

18—ACIDS, ALKALIES, SALTS AND SUNDRIES

E. M. SYMMES

The concentration of weak nitric acid. G. PAYRAS. *Rev. gén. mat. plastiques* 6, 147-51 (1930).—A description is given of the Chem. Construction Co.'s Fr. pat. 536,108 (C. A. 23, 4026). A. PAFINEAU COUTURE

Vanadium as a catalyst for sulfuric acid manufacture in America. ENSIO ALHO-PURO. *Acta Chemica Fennica* 3, 112-4 (1930).—The characteristics of V catalysts now used in plant operation are given. The best results are obtained by the Selden V mass and the new converter developed by A. O. Jaeger, which give as high as 98% conversion. In present day plant operation V gives a higher av. conversion than Pt, because Pt is contaminated with traces of As and Cl. The price of Pt is 184-253 times the price of V, but the high return value of Pt catalysts and the high price of licenses for V catalysts make this proportion lower. H_2SO_4 plants which were operated or built in 1929 and which use the Selden converter system have a total capacity of more than 300,000 tons of H_2SO_4 per year. The total production of H_2SO_4 by the contact process in the U. S. during the same year was 1,870,000 tons as 100% acid.

S. A. KARJALA

Potash. GEO. W. STOCKING. *Chem. Markets* 28, 247-52 (1931).—The industry in Germany is described. E. H.

A study of the properties of polyhalite pertaining to the extraction of potash. VI. Experiments on the production of potassium chloride by the evaporation of leach liquors from decomposition of uncalcined polyhalite by boiling saturated sodium chloride solutions. H. H. STORCH and F. FRAAS. *Bur. Mines. Rept. of Investigations* 3062, 7 pp (1931), cf. C. A. 25, 1338.—Evapn. of 90% of the H_2O , interspersed with 3 crystn. steps, will yield 73% of the K_2O as crude KCl, the bulk of the impurities being NaCl. This product may be refined readily by recrystn. to produce pure KCl. A preliminary estimate indicates about \$20 a ton as the cost of production at the plant. A. H. E.

The acids and acid salts which contaminate cream of tartar. FILIPPO PERCIABOSCO. *Atti III congresso naz. chim. pura applicata* 1930, 520-5.—Some cream of tartar contains appreciable quantities of $Ca(C_2H_3O_4)_2$ instead of KCl. E. M. S.

Quinquennial review of the mineral production of India for the years 1924 to 1928. Phosphates. E. H. PASCOE. *Records Geol. Survey India* 64, 413-21 (1930).—Apatite is known in many parts of India, some deposits assaying 20-25% P_2O_5 . Nodular deposits of lime phosphate also occur.

ALDEN H. EMERY

A brief history of the world's phosphate rock production. A. N. GRAY. *Superphosphate* 4, 25-46 (1931).—Historical notes relating to the discovery of phosphate deposits and the early production of phosphate rock in all parts of the world are given. The world's annual production of phosphate rock by countries is given for the years 1847 to 1929. K. D. JACOB

Dehydration of salts of phosphoric acid. S. S. DRAGUNOV. *Udobrenie i Urozhas (Fertilizers and Crops)* 2, 409-16 (1930).—Graphs show the transformations of the salts of Na, K and Ca or phosphoric acid upon drying. From these it is possible to judge the % of pyro- and meta-modifications formed, as well as the speed of the transformations at various temps. With the loss of H_2O of crystn. $CaHPO_4 \cdot 2H_2O$ becomes less sol. in citrate soln. At 100° only 4% of the H_2O of crystn. was lost in 1.5 hrs. At 150° 19% was lost in 20 min. It is important not to dry phosphate very much above 100° for any period of time (not more than 30-50 min. at 100°). Ca metaphosphate, insol. in acid, can be hydrated in the autoclave at high pressures and brought back to the sol. state. J. S. JOFFE

Utilization of the gases obtained in the process of volatilizing phosphorus and the production of soluble phosphates. A. P. DUNAIEV. *Udobrenie i Urozhas (Fertilizers and Crops)* 2, 397-409 (1930).—The relation of H_2O and O from the air to the system: P, PH_3 , CO, H, in which a reaction of selective oxidation of generator gas takes place was investigated. A modification of the app. used by Britzke and Pestov (cf. C. A. 24, 208) was necessary. A description and drawing of the new app. are given. The exptl. data obtained show that the oxidation of P with H_2O over CaO at a temp. from 550° to 950° is completed to phosphoric acid at fairly rapid velocity of the gas stream. Very little PH_3 is found in the gases, and the sum of CO + H is not decreased. The materials used for oxidation contained no phosphates. In the oxidation of the gases with air no excess of O was necessary in the process, except what was needed for the oxidation of P. At the lower limits of temp. no O is used up by the CO. The generator gases used in this manner can be used again for heating or other purposes. After

NUMBER 7

CHEMICAL ABSTRACTS

〇 ≈ ≈ ≈ 〇

PUBLISHED SEMI-MONTHLY BY
THE AMERICAN CHEMICAL SOCIETY



APRIL 10, 1931

ABSTRACTORS

C. A. Abieda	W. C. Pfaugh	V. Katschewsky	Withrow Mores	H. G. Schurecht
J. H. Adams	M. A. Mady	K. Kambe	H. H. Mosher	Genther Schwach
C. R. Addison	A. L. Elder	A. L. Kammherer	A. A. Madsen	Janet D. Scott
W. H. Adolph	P. H. Emmert	K. Kashima	G. W. Mahlemas	W. Segerblom
Jerome Alexander	K. H. Engel	Louis Kelley	C. E. Matlin	Florence Seibert
D. K. Alger	J. B. Patrick	R. W. Kelly	A. M. Murphy	M. R. Seymour
P. O. Anderson	Alice W. Poperson	O. L. Kelso	J. T. Myers	P. L. Seymour
B. C. Anderson	R. C. Erb	J. A. Kennedy	P. J. Nakamura	Jones
R. L. R. Ardagh	Paul Fuchs	A. L. Kibler	Leo Namasovich	D. E. Sharp
L. K. Arnold	J. L. Evers	C. C. Kline	N. M. Naylor	R. M. Saelton
J. B. Austin	S. L. H. Eibertson	V. Kihara	E. Nelson	A. P. Shepard
Janet F. Austin	G. M. Keana	C. G. King	O. A. Nelson	O. E. Shappard
F. P. Badger	R. H. Ewell	P. Klamms	P. T. Newsum	V. Sheshanoff
C. H. Badger	Mary A. Laver	H. S. van Kloostm	J. H. Nicot	K. Shunkley
V. E. Bakrow	L. T. Fuchait	S. Kouda	J. G. Norderoore	J. W. Shipley
C. V. Bailey	C. R. Fellers	M. V. Lomdady	I. H. Odell, Jr	H. B. Sims
L. H. Bailey	A. T. Fellows	I. A. Kover	W. Ogawa	P. A. Simmonds
A. Ballard	R. P. Ferguson	Frances Krumm	H. L. Olin	D. E. Singer
J. Balozas	R. H. Ferguson	H. R. Kratbill	Takeshi Oniki	A. S. Smith
R. C. Bardwell	W. C. Ferrelino	H. P. Kruege	Lara Onager	O. M. Smith
R. M. Barnette	S. S. de Fendly	Antonina Lefers	P. H. Overall	Corneilus T. Snell
A. E. Barthel	Arthur Pleischer	S. Kacimawa	W. L. Owen	P. D. Soell
P. H. Bergum	Stephen K. Ford	J. M. Lachon	C. W. Owens	J. E. Snyder
R. Beutner	A. B. Foster	P. W. Laird	A. Paparaz	K. Soderberg
E. Beckum	D. K. French	R. H. Lambert	Costara	C. W. Sooders
F. B. Billington	W. G. Lewerlet	H. A. Lahan	H. C. Fench	A. Soule
C. B. Blomham	E. Gebauer	W. D. Langley	J. L. Parsons	M. H. Soule
Charles Blons	S. Fennberg	S. Lander	I. J. Patton	b. C. Sorrentoff
A. A. Bochiuna	F. G. Geremuth	V. G. Lata	C. H. Peet	A. D. Stillman
W. Bojars	L. E. Gilson	H. W. Lecky	F. R. Prikshaning	H. M. Stark
W. H. Burston	William Gilson	Stacy H. Lear	W. A. Jersberg	H. W. Stiegler
L. B. Brock	G. R. Greenough	J. P. Leete	L. Fessel	W. W. Stiller
A. M. Brant	F. R. Greenbaum	L. J. G. de Loran	C. E. Peterson	H. Steier
C. A. Brautrecht	P. P. Grublar	W. G. Lighthouse	L. M. Petty	B. B. Stone
G. A. Bruns	A. Grutman	Sam Leuber	A. J. Phillips	J. R. Strong
Egon Brutscher	F. C. Haba	H. P. Leopold	E. Pickering	O. G. Sward
J. P. Brewster	L. P. Hall	Amy Lebowitz	Yetta Polonowsky	A. Swaid
J. Brockway	Benjamin Hamilton	B. S. Lerner	L. L. Powers	P. W. Tanser
W. R. Brode	V. H. Hartman	S. M. Levine	L. A. Prudson	S. Tashiro
C. L. Brodie	Milton Harris	Edward Lawie	M. J. Prentiss	A. L. Taylor
Doe Broome	B. Harrow	J. H. Lewis	L. L. Quill	D. B. Taylor
D. P. Brown	W. D. Hatfield	R. H. Lombard	Edith Quimby	W. C. Taylor
F. E. Brown	T. G. Hawkey Jr	E. R. Long	O. T. Quimby	R. W. Thiele
J. B. Brown	R. A. Herold	W. R. Loomis	T. R. Ragan	K. V. Thimann
Rachel F. Brown	A. L. Keane	C. H. Long	L. S. Ramadoff	W. Thomas
B. C. Brunstetter	J. S. Hepburn	A. P. Lockrop	P. H. Rathmar	P. A. Thumamert
R. E. Bullock	K. L. Herber	O. H. W. Lucas	Albert L. Rawlins	R. H. Thompson
M. D. Buschmeyer	A. C. Higgins	J. P. Lyman	C. L. Read	D. Thunau
R. C. Butall	C. A. Hill	C. M. McCay	J. L. Reedy	W. H. Tiffany
A. Batta	John R. Hild	M. McMillan	L. H. Reperson	G. Twardy
L. W. Bels	Wm. L. Hill	J. C. McNally	D. A. Keyokids	J. G. Tolpin
J. S. Caldwell	Ann H. Hird	M. W. McPherson	C. H. Richardson	A. P. H. Treutill
G. Calogheri	Alfred Hirsch	A. R. Macommac	T. H. Rider	P. H. Untiedt
A. T. Cameron	A. B. Hitchcock	S. A. Madorsky	Cecilia Riegel	Frank Urban
F. Camps Campese	R. C. Hockats	K. H. Mann	W. Kiding	R. L. Vander
T. M. Carpenter	W. W. Hodges	Ken Manzel	C. A. Kobak	Bochs
R. H. Carver	B. J. C. van der	C. L. Mantrill	K. J. Roberts	N. Van Patten
T. H. Chilton	Hoover	Frank Marshak	L. D. Roberts	W. K. Vassagan
R. K. Clark	A. W. Holmes	P. W. Marsh	C. C. Roberts	V. V. Vassilowsky
R. W. Clark	H. P. Holmes	Leona Marsh	Jean Rogers	A. J. Vorneld
C. L. Clark	E. S. Hopkins	C. W. Maslow	B. R. Roethli	Louis Waldman
J. D. Clarke	Robert Howatson	Peter Mazzetti	M. C. Rogers	S. Waldbott
A. W. Conner	C. E. Hrobenky	G. E. Matthews	A. R. Rose	H. W. Wether
H. J. Corper	C. J. Humphreys	Marie Miesin	W. Gordon Rose	R. S. Welles
Emma R. Crandall	P. M. Humphreys	L. A. Maynard	W. H. Rowe	P. D. Rowen
B. N. Dandoli	E. Horvath	W. F. Mayers	Gerhard Rahn	L. E. Warren
J. D. Davis	W. J. Hua	R. F. Mehl	C. C. Ruckhoff	P. J. Weber
T. L. Davis	Harbert Isdey	E. G. Henter	R. J. Ruck	H. B. Weiser
Paul R. Dawson	T. Ince	H. B. Merrill	R. W. Ryan	Wm. West
R. S. Deane	C. E. Isee	A. R. Stern	A. P. Sachs	Abraham White
J. A. Denison	F. V. Jackson	A. E. Meyer	R. Salto	Annette Wirben
W. L. Denman	K. D. Jacob	B. Miller	H. K. Salatsberg	Jan Wiertelak
H. J. Denel Jr	L. Jacoby	H. C. Miller	E. B. Sangar	J. J. Williams
A. H. Dierker	C. L. F. Jefferys	L. B. Miller	R. Sassner	R. C. Wilson
D. B. Dill	M. A. Jerzak	J. P. Miller	R. Schaud	C. L. Wilson
Malcolm Dols	J. S. Joffe	A. L. Mitchell	Dawson Schaaf	C. W. Wilson
R. H. Doughty	A. H. Johnson	A. E. Min	E. R. Schaler	Raymond W. Ioss
A. W. Dix	A. K. Johnson	A. J. Monack	E. R. Scherer	Chas. Wirth, III
Arne Dröghem	H. F. Johnston	B. Moorhead	Robert Schwab	Ruby K. Warner
C. F. van Dura	C. L. Jones	H. R. Moore	C. J. Schillenberger	G. R. Vobe
A. G. Dulmet	S. Jozsa	J. H. Moore	Florence N. Schott	C. Zachin
P. L. Duna	W. Jung	S. Morgulis	E. Schotts	F. W. Zerbao
H. C. Foss	J. C. Jurek			
Harry Ezrie				

Abstracts signed B. C. A. have been obtained by special cooperative arrangement with British Chemical Abstracts

CHEMICAL ABSTRACTS

Editor: E. J. Crane, Ohio State University, Columbus, Ohio.
Associate Editors: Elmer Hookett, Hylan Game and Leonard T. Capell
Assistant Editors

M. S. Anderson	D. J. Demarest	P. R. Hawk	L. A. Olney	J. J. Skinner
W. L. Badger	Allen H. Emery	P. H. Howe	A. M. Patterson	W. R. Stearns
G. R. Barton	W. H. Emery	C. H. Kerr	T. O. Phillips	C. O. Storm
Edward Bartow	W. P. Parasher	H. A. Lepper	A. N. Richards	E. M. Symmes
R. Benedict	A. C. Fieldner	R. K. Marshall Jr	Richard Rimbach	P. P. Underhill
P. C. Black	C. G. Pink	A. R. Middleton	Allen Rogers	H. O. Welle
F. L. Brønne	C. N. Frey	Harlan S. Miner	C. A. Rouiller	C. J. West
C. H. Curran	H. W. Gillett	Chas. B. Morrey	A. H. Sablin	K. T. Wherry
I. K. Dale	H. A. Goerner	C. E. Monroe	J. P. Schairer	H. P. Wahlman
C. C. Davis	W. T. Hall	W. A. Noyes Jr	E. Schenckel	J. C. Witt

CONTENTS

Patents	I	16. The Fermentation Industries	1525
Author Index	II	17. Pharmaceutical Chemistry	1630
1. Apparatus and Plant Equipment	1615	12. Acids, Alkalies, Salts and Sun-	1610
2. General and Physical Chemistry	1617	dries	
3. Subatomic Phenomena and Radio-	1614	19. Glass, Clay Products, Refractories	1647
chemistry		and Enamelled Metals	
4. Electrochemistry	1641	20. Cement and Other Building Ma-	1661
5. Photography	1649	terials	
6. Inorganic Chemistry	1655	21. Fuels, Gas, Tar and Coke	1654
7. Analytical Chemistry	1656	22. Petroleum, Lubricants, Asphalt	1662
8. Mineralogical and Geological	1669	and Wood Products	
Chemistry		23. Celluloses and Paper	1677
9. Metallurgy and Metallography	1672	24. Explosives and Explosions	1676
10. Organic Chemistry	1682	25. Dyes and Textile Chemistry	1678
11. Biological Chemistry	1680	26. Paints, Varnishes and Resins	1683
12. Foods	1683	27. Fats, Fatty Oils, Waxes and Soaps	1687
13. General Industrial Chemistry	1693	28. Sugar, Starch and Gums	1697
14. Water, Sewage and Sanitation	1696	29. Leather and Oils	1702
15. Soils, Fertilizers and Agricultural	1611	30. Rubber and Allied Substances	1706
Poisons			

PATENTS

A statement giving information as to how to obtain patent specifications and drawings, both United States and foreign, is to be found on page I of *Chemical Abstracts*, 24, No. 5.

Published by THE AMERICAN CHEMICAL SOCIETY, which also publishes *The Journal of the American Chemical Society and Industrial and Engineering Chemistry*.

Communications for the editor should be addressed to E. J. Crane, Ohio State University, Columbus, Ohio.

Remittances for annual dues (\$15.00) in the American Chemical Society or for subscriptions to this *Journal* (\$7.50 per year), changes of address, orders for subscriptions or back numbers, claims for missing numbers and all matters intended for the *Proceedings* should be sent to the Secretary, Charles L. Parsons, Mills Bldg., Washington, D. C. Foreign postage is extra, being \$4.50 per year added to the dues of members residing outside the United States, its possessions and the countries of the Pan-American Union and \$1.50 per year for *Chemical Abstracts* alone, except that only one-half of these amounts is added for journals going to Canada. The countries of the Pan-American Union are Argentina, Bolivia, Brasil, Chile, Colombia, Costa Rica, Cuba, Dominican Republic, Ecuador, Guatemala, Mexico, Nicaragua, Paraguay, Peru, Republic of Honduras, Republic of Panama, Salvador, Spain and Uruguay.

The Council has voted that no claims will be allowed for copies of journals lost in the mails unless such claims are received within sixty days of the date of issue, and no claims will be allowed for issues lost as a result of insufficient notice of change of address. "Missing from files" cannot be accepted as the reason for honoring a claim—Charles L. Parsons, Secretary.

Published by the American Chemical Society, Easton, Pa. Secretary's Office, Mills Bldg., Washington, D. C. Entered as second-class matter at the Post Office, Easton, Pa. Acceptance for mailing at special rate of postage provided for in Section 1103, Act of November 3, 1917, authorized July 13, 1918. Mack Printing Company, Easton, Pa.

AUTHOR INDEX

[An asterisk (*) indicates that the reference is to an abstract of a patent]

- Alaudie P 1410
 Albright O D 1354
 Alderhailen I 1316
 Ale M 1554
 Aleku I 1564
 Aichard C 1573 1573
 Ackert J F 1559
 *AC Spark Plug Co 1416
 Adair P F 1540
 Adam N K 1422
 Adams C I 1467
 Addons R M 1512
 Agafonoff M V 1611
 van Aegelen C 1623
 Aginteev K K 1435
 Aharoni J 1440
 Ahlfield F 1467
 Ahmad B 1560
 Ahmann C F 1596
 Ahrens W 1470
 *Akt Ges Brown Noveri & Cie 1416
 Alexander E 1471
 Alexander S G 1631
 *Algemeent Kunstnijde Uni N V 1673
 Alverti G 1413
 Allam P 1423
 Allan J 1633
 Allemayer 1616
 Allibone T P 1439
 Almia A 1477
 Alberg C L 1543
 Almsa F D 1571
 Altpeter J 1505
 *Aluminiums Bergbau und Industrie A G 1653
 Alsous L 1577
 *Amalgamated Carbutters Ltd 1667
 Ambler H R 1413
 Amen C C 1622
 *American Machine & Foundry Co 1481
 *Ammoniaque chimique et dérivés soc anon 1643
 Amshlee J W 1509
 Andreau, W 1681
 Anderson A L 1463
 Andreev K K 1674
 Andrews, A I 1650
 *Anke A 1651
 Anker E 1638
 *Anode Rubber Co Ltd The 1703
 Ansbacher S 1548
 Anselmino K J 1564
 Anton G 1582
 Aoyama S 1533
 Applebey, M P 1424
 Arakawa T 1535
 Arcand A 1578
 Arkhangel A D 1467
 *Arkless, P 1415
 Arnold H 1536
 Aery, A C 1673
 Aron M 1547
 Asahina T 1509
 Asai T 1449
 Asaka K 1651
 Ascham L 1537
 Ascham O 1451
 *Associated Portland Cement Manufacturers Ltd 1633
 Aston B C 1602
 Aston F W 1419
 1437
 Atton G H 1436
 Attree G P 1615
 Aubert H 1571 1573
 Auer L 1646
 Auerbach J 1515
 Auerbach L 1603
 *Auerbach R 1416
 *Aug Nowack A G 1645
 Aulich M 1433
 Aumeras 1440
 *Auren G 1651
 *Babcock & Wilcox, Ltd 1661
 Bachér P 1515
 Bachmann W 1572
 Back E A 1679
 Backus H 1637
 Bacon R F 1641
 *Bader, W 1415
 Bagg H J 1580
 Baggegaard-Rasmussen H 1636
 Bailey, C H 1506
 Bailey, K C 1431
 Bain, E C 1499
 *Bakelite Corp 1651, 1692
 Baker G L 1601
 Baker, N W 1577
 Bakker, C J 1438
 Balini, J 1589
 Balis A K 1542
 1543
 Baneroff G 1562
 Bancroft W D 1426, 1562 1590
 Bandy M C 1460
 Bannar E 1599
 Bansen H 1473
 Barbella N G 1639
 Barthier, G A 1454
 Barger, G 1632
 Barkun, G 1557
 Baroch, C T 1443
 Barraud, M (Miss), 1665
 Barrensheen H K 1514 1574
 Barreite, C S 1479
 Barrett P F 1441
 Barnes D L 1649
 Barthelmy, H L 1679
 Bartels W B 1475
 Barth T 1459
 Bartsch O 1614
 Bary P 1426, 1676
 Basart J C M 1421
 Basset, J 1472
 Bass U 1525
 Bateman A M 1464
 Bateman G C 1641
 Bateman F R C 1639
 Bauer, K. H 1504
 Bauer, L 1562
 Bauer, O 1479
 Bauer W 1563
 *Baugh & Sons Co 1645
 *Baum E 1537
 *Baumann A W 1603
 *Baumann, P 1643
 Baumgarten T 1615
 Baumgartner, W H 1609
 *Bayerische Stickstoffwerke A G 1612
 Beach A C G 1645
 *Beaudouin L 1706
 Becken E 1569
 Bechard M 1609
 *Bechert, B 1451
 Beck G 1436
 Beck, W 1478
 Becker, W 1470
 Beckmann, P E 1638
 *Bedwell H E 1653
 Beiles I 1587
 Belcher C P 1469
 Belyaev, N 1631
 Bemmann, R 1432
 *Bendlin R 1606
 Benetato G 1554
 Benetato M 1587
 Benn G S 1702
 Bennett D 1589
 Bennet, H 1474
 Bennett, C T 1634
 Bennett, T I 1579
 *Benoit, L 1603
 Benson, L H 1459
 Bernard, H L 1597
 Berend N 1598
 Berger, R 1544
 Berget A 1441
 Bergmann, M 1510
 Berkley 1427
 Berknor F 1595
 Berl F 1432
 Berlin, L E 1619
 Berliner, E 1605
 Berliner, J F T 1610
 Bernan H 1459
 Bernard J I 1441
 Bernardino C 1434
 *Bernstein A 1672
 *Berry, R E 1674
 Bertram K 1556
 Besborodov, M A 1667
 Beyer, O 1636
 von Bezdold 1414
 Bhattacharya S N 1455
 Bhattacharyya, T 1530
 *Biacco, G 1415
 Bigwood R J 1425
 Bijvoet, J M 1435
 *Bilster, J 1444
 Binet, L 1571
 Bingham, E C 1419
 Bignon, O R 1579
 Birkenbach, L 1454
 Birkhaug K P 1621
 Bischoff, P 1645, 1569
 Biehopp, F C 1621
 Bism, M 1492
 *Bitumula Kalkaphalt A G 1667
 Buzzell J A 1613
 Björjyke, H 1462
 Blackett P M S 1438
 Blake M A 1552
 Blanchard, R 1463
 Blanchetiere, A 1571
 Black E 1471
 Blaszkowska, Z 1419
 *Bleasdale H 1671
 Blier, P 1500
 *Blie C J 1697
 Blood, P T 1551
 Blum, L 1574
 Blumano, A 1513
 Blum Bergmann, O 1517
 Blumenthal, A 1414
 Bocquenois 1585
 Bocquenois, A 1585

- Bode H 1656
 Bodewig I 1674
 *Böhme, H T A G 1643
 Boepfling 1450
 *Boccone P 1642
 Bohn J I 1137
 Bohn N H H 1441
 Bohrer H G 1558
 Bowlin A 1555 1562
 Bolam T R 1623
 Bolin P M 1599
 *Bollmann H 1656 1701
 Bolsover, G R, 1475
 Bole, P, 1496
 Bond P A, 1422 1427
 Bonino G B 1477
 Bonner, L, 1674
 Bonström J W, 1417
 Borli, 1583
 Borlet, P, 1594
 Borghi, R, 1569
 Borgmann K, 1441
 de Borgssof L, 1581
 Borsche W, 1510
 Borsook, H, 1567
 Borsi, M, 1567
 *Borsten H, 1605
 *Borzykowski, B, 1672
 Bosch C, 1434
 Bose, P K, 1532
 Bosnjak, D, 1412
 Bost, R W, 1522
 Boswell, P P, 1463
 Botolffsen, R 1492
 *Bouchet, J, 1671
 Boulanger, P, 1578
 Boularie, A, 1473
 Bowen, A R, 1661
 Bowley, J L, 1656
 Bowles, O, 1641
 Boycott, A B, 1579
 Boynton H W, 1641
 Bracconnot, J, 1674
 Bradshaw, J H B, 1650
 Bradway, R M, 1561
 Brady, G A, 1657
 Brady, I P, 1462
 Brady, O L, 1506
 *Braddon, R I, 1691
 Braknuchast, H N, 1530
 Brambilla, S, 1560
 Bramdenburger, H, 1677
 *Brandt, D G, 1666, 1667
 *Brandt, R, 1661
 Brantley, J C, Jr, 1634
 Brauch H D 1658
 *Brassett, H A, 1413
 Bratter, H M, 1641
 Brauman P 1454
 Bray, W C 1430
 *Bregl J H 1662
 Breith I 1533
 *Briemer, P 4616
 *Briempe W I 1644
 Brietsmajles S 1429
 Brice A T Jr 1518 1577
 Brief M 1547 1549
 *Briggs E H 1413
 Bright R D 1623
 Bright H A 1456
 Brinkman R 1548
 *Brinkmann T 1391
 *British Celanese, Ltd 1645 1667
 *British Research Association for the Woolen & Worsted Industries 1641 1644
 Britke, E V 1470
 *Britke, E V, 1641
 Bronn, J, 1644
 Brnoka, M M, 1552
 Brown, R B, 1620
 *Brown, C, 1667
 Brown, H R, 1675
 Brown, R H, 1478
 *Brown, S, 1416
 Browne, P L, 1685
 Bruce, E L, 1163
 Brühl, P, 1475
 Brün, W, 1446
 Brühler, P, 1705
 Brulins, G, 1437
 Brunl G, 1420
 Brus, G, 1512
 *Brusac, J I A, 1666
 Bryan, O C, 1614
 Bryant, S A, 1522
 Buchan, S, 1501
 Buchner, E H, 1426
 Buckley, H E, 1420
 Bulon, T, 1634
 Budnikow, P P, 2049
 Büchl, J, 1634
 Bulow, M, 1511
 Bürgl, P, 1674
 Bugnani, L, 1566, 1567, 1573
 Bulson, N (Mile I, 1554
 *Bullard Co., The, 1444
 Dutchard, E F, 1466
 Burfoot, J D, Jr, 1467
 Burgess, W. G, 1421
 Burgess, G K, 1630
 Busk D, 1512
 Burselmer W 1491
 Burns G I 1601
 Burns, K (Miss) 1479
 Buscatti I H 1521
 Buschmann A 1554
 Butcher C H 1611
 Butomachon H 1615
 Butowski W 1611
 Byers H C, 1612
 Baler R R 1472
 *Cable D A 1651
 *Cahot S 1671
 Cachera H 1587
 Cahane M 1571
 Cajala H 1636 1637
 Caley I R 1456
 *Cakoff W S 1706
 Callidine I 1677
 Cameron A T 1591
 Cameron G H 1579
 Campbell A D 1563
 Campbell I A 1564
 Campbell M R 1469
 *Canadian Electro Products Co Ltd 1692
 *Canning, J R, 1413
 Cantelow, A 1561
 Cantegat I I, 1574
 *Carbolic Co., The, 1641
 Carroll, G, 1462
 Carrozo J O, 1702
 *Carter, R S, 1416
 *Casale, I, 1612
 Casazza I, 1453
 Cadmie, H, 1476
 Casparis, P, 1637
 Castelsin, I, 1426
 Caulotte, P, 1570
 Caulkins A L, 1430
 Cavala, 1604
 *Ceba Ltd, 1447
 *Cellulose Corp, 1672
 Celd, S A, 1636
 Cercher, V, 1493
 Chabrol R, 1455
 Chahoeitch X, 1556
 Chambers, R, 1504, 1591
 Champion, P C, 1454
 Chanoir, M, 1427
 Chapas, G, 1641
 *Chapelle, P A, 1697
 Chaplet, A, 1601
 Chappuis, J, 1441
 Charmetant, C, 1611
 Charonnat, R, 1452, 1543
 Chartier, J, 1676
 Chase W W, 1677
 Chelova, T N, 1701
 *Chem Fab von Heyden A G 1605 1626
 *Chem Fab Ludwig Meyer 1626
 *Chemische Fabrik vom Sandt 1681
 *Chemische Industrie von Hasselt 1603
 Chertshier R 1548
 Cheselievik V A 1702
 *Chester R H 1614
 Chetelins A 1577
 Chigorev D A 1621
 Child R 1570
 Childs A A 1617
 Childs I 1622
 *Chilowsky C 1662
 Christie A W 1600
 Christmann A A 1544
 Chrusak I 1436
 Chroster T 1551
 Chulb, H M 1674
 Chutshov I 1673
 Chwala, A 1621
 Chzapelo S 1581
 Chwars, A, 1465
 *Clapp A I 1616
 *Clariana A I 1444
 Clark A T 1606
 Clark I P 1591
 Clark, R S 1622
 Clarke G G 1457
 Clarny J, 1678
 Claude H, 1544
 Clayton, R T, 1442
 Clemm, R R 1570
 Crennell, J B, 1457
 de Clerck, J, 1624
 Cliche H 1604
 Clirel, L, 1427
 Coal Tryor, R A, 1644
 Cocking T T, 1631
 Columbus A, 1575
 Coehn, A, 1454
 Coelhu R, 1591
 Coffey, H, 1507
 Cohen, H 1701
 Coleman, D A, 1596
 Coleman, J M, 1550
 Colin, H, 1553
 Collet, P, 1546
 *Colle J H, 1414
 Collin, I M, 1457
 Collip, J B, 1563
 *Compagne, R, 1450
 *Compagnie Française d'accumulateurs électriques, 1645
 *Compagnie Française

- l exploitation des procédés Thomson Houston 1640, 1644
 *Compagnie des mines de Vicoigne, Noeux et Drocoourt, 1660
 *Compagnie nationale de matières colorantes et manufactures de produits chim. du Nord Réunies (Établissements Kuhlmann) 1679
 *Compagnies réunies de gaz et d'électricité 1662
 *Compteur des textiles artificiels, 1672
 Commie A. A. D., 1623
 Coma, M. W., 1522
 Conover J. D., 1641
 Consortium für ein trocknem Ind. G. m. b. H., 1537 1692
 *Costal, C., 1451
 Cook A. S., 1592
 Cook S. J., 1641
 *Cooley A. C., 1451
 Coons, A. T., 1641
 Copp H. F., 1637
 *Coppée É., 1642
 Corbellini A., 1529
 Corbino C. L., 1593
 Cornelius H. P., 1461
 *Cornillot A., 1704
 Correa, C. W., 1470
 Cosmulesco I., 1571
 Cosmulesco J., 1573
 da Costa, S. F. G., 1553
 Cotton R. T., 1679
 *Cotton, R. T., 1626
 *Coulter S., 1644
 Courtois, A. (Silla), 1603
 Coultis J. R. H., 1613
 Cowperthwaite, L. A., 1455
 Craig D. N., 1433
 Craig L. C., 1624
 Cramer, C., 1606
 Cramer, H. R., 1606
 Crandall, P. R., 1413
 Crane, C. H., 1641
 Crawford N. Z., 1558
 Crisp, A., 1677
 Cristiani, H., 1586
 Cristol, P., 1574
 *Croft C. M., 1686
 Crowe J., 1423
 Cumpe Z., 1633
 Cummings, M. J., 1557
 Curtis O. F., 1550
 Curtman L. J., 1459
 Curtis, R. M., 1477
 Cutanda, V., 1621
 Dadies J., 1571
 Dalfert O., 1637
 *Dalgleish, H. V., 1416
 *Dallas, J. C., 1611
 von Dallwitz Wagner R., 1417
 von Dallwitz Wagner R., 1415
 Dalmán J. M. Pia 1434
 Dalmeze, J., 1546
 Dalström, Y., 1476
 Damm, P., 1562
 Dane, E., 1533
 *Dangle, D. H., 1416
 *Darleg J. P., 1643
 *Darlington, H. J., 1451
 Dertens, G., 1505
 Desbatsville, L. J., 1451
 *Desbatsville M. J., 1673, 1657
 Davey W. P., 1423
 Davidson J. G., 1671
 Davidson W. M., 1621
 Davies, C., 1414
 Davies G., 1453
 Davies, J. S. H., 1513
 *Davis P. W., 1451
 *Debuch C. P., 1490
 Debusigne J., 1603
 Decker S. W., 1551
 Dedek J., 1700
 Deem, J. W., 1673
 *Dejays, J. J., 1450
 Delbart, G., 1473
 Delite, P., 1566
 Demay, A., 1463
 *Dempsier, R. & J., Ltd., 1667
 Demsch, 1616, 1619
 De Fasnus, L., 1438, 1504
 Detrens E., 1574
 *Deschauer, A., 4536
 Deubel F., 1461
 *Deutsch H., 1537
 Deutsche Abwasser Reinigungs-Ges. m. b. H., Schädlerengung 1614
 *Deutsche Gold- und Silber Schendens stalt vorm Roessler 1539 1643 4651
 Deutscher K., 1505
 De Vasconcelles, 1592
 Devoto J. S., 1649
 Dhar, N. R., 1540
 Dhéré, C., 1410
 Diacout, K., 1510
 DiBella F. A., 1602
 *Dodet, L. J. B., 1451
 Dicho, G., 1700
 *Dietzle E. A., 1661
 Dillinger, H. L., 1444
 Dittley, W., 1513
 Dumbleby, T. E., 1644
 Dumbleby, V., 1647
 Demolese A., 1584
 Dinger, K., 1667
 *Diablerie M. A. A. G., 1490
 Diehlage, R., 1513
 Dieche L., 1566
 Dittler E., 1461
 Ditz, H., 1455
 Dixon, S. M., 1632
 Djuricic I., 1571
 Dm. thov, N., 1694
 Dobbs, F. J., 1442
 Dobrovol'skaya, O. M., 1616
 Dodds E. C., 1576
 Dodge, W. R., 1472
 *Dohet, E., 1662
 Dohack A., 1699
 Domashewicz, M., 1676
 Donaldson J. W., 1473
 Donovon P. P., 1494
 1499
 Dopf, K., 1631
 *Dorr Co., The 1415
 *Douglass, W. A., 1706
 Drake E. B., 1477
 *Dreaper, W. P., 1673
 *Dreyfus, H., 1536, 1537, 1696
 *Dreum J. J., 1447
 Drachman, D. V., 1615
 Dvorskuvich, A. A., 4674
 *Duboss, G., 1640
 Dubois J., 1655 1658
 Dubois P. A., 1495
 Dubsky, J. V., 1414
 *Dudzere Corp. of America, 1451
 Durr, W., 1496, 1497
 Duff D. C. B., 1413
 *Duffield F. L., 1661
 Duhaie B. M., 1704
 Duhaie H. H., 1663
 Dumanski, A. V., 1423
 DuMex A. G., 1634
 Dunbar, R., 1417
 Dunham R. S., 1623
 Dunlap H. L., 1457
 Dunlop D. M., 1555
 *Dunlop Rubber Co., Ltd., 1674 1703 1706
 Dunn, S., 1551
 Dunstan E. T., 1442
 Dutton A. R., 1604
 Dupasay, R., 1422
 *Duplate Corp., 1651
 *Du Pont de Nemours & Co., 1537, 1643 1657, 1692
 Durand, J. F., 1401
 Durno F., 1497
 Durtat I., 1674 1679
 Du Toit, M. S., 1615
 Durti, A., 1630
 Dutton W. C., 1622
 Dural, M., 1597 1647
 Dyes W. A., 1677
 Dyke, W. J. C., 1445
 *Dynamit & Co. vorm. Alfred Nobel & Co., 1675
 *Dyryak, V., 1653
 Dyson, G. M., 1603
 Dzwoski, K., 1503, 1514, 1515, 1519
 *Eaton, H. D., 1416
 Eaton, W. C., 1657
 Fbel A., 1577
 Fierhard, G., 1441
 Ebert, L., 1497
 Eckel, E. C., 1641
 Eckel, C. F., 1417
 Eder, R., 1635
 *Eisen und Stahlwerk Hoesch & Co., 1416 1625
 Etel, W., 1471
 Fierfors H., 1586
 Ekkert, L., 1633
 Ekman, W., 1476
 Elgin, J. C., 1663
 Ellis, C., 1536
 Ellis, C. D., 1436
 *Ellis G. H., 1657
 Fmerqua L. (Mile), 4561
 Emole, B. L., 1616
 Engel H., 1677
 Engel P., 1558
 *Engelbrektion, O., 1654
 Engelder, C. J., 1459
 Engelhardt, V., 1447
 Engela, O., 1616 1620
 *Engels, W. H., 1629
 Erdmannsdorffer, O. H., 1461

- Ernst P. 1567
 Ertel, L. 1535
 Eschenbrenner, H. 1631
 Espach R. H. 1663
 v. Euler H. 1516, 1554
 Evans, H. M. 1526
 Evans R. P. 1602
 Evans U. R. 1414
 Evers, G. V. 1645
 Evvard J. M. 1545
 Ewles, J. 1434

 Fabre R. 1591
 Fabry, C., 4415
 Fachgruppe für Wasserchemie d. Vereins deutscher Chemiker 1610
 *Fahrenwald P. A. 1450
 Fairchild, I. J. 1650
 Falta W. 1564
 *Farrall Products Co., Inc. 1151
 Farkas, G. 1563
 Farmer, E. H. 1432
 Farquharson, D. J. 1642
 Farrar, M. D. 1623
 Fau, J. 1545
 Favier, H. 1637
 Fawcett, G. 1577
 Faust, K., 1595, 1632
 Feistel, F. 1502
 Fell, H. B. 1547
 Felt, W. L. 1657
 *Feltz & Gussensmum Catlawerk, A. G. 1419
 Fenner, C. N. 1464
 Fenz, E. 1545
 Ferguson, J. 1424
 *Fernbach, A. 1630
 Ferrari, A., 1420
 Feustel, I. C. 1612
 *Fiberloid Corp., The, 1672
 Fielder, F. D. 1447
 Fielding, W. L. 1602
 Fikentscher, R. 1576
 Filon, L. N. G. 1441
 Filosofov, M. S. 1698
 Findlay, G. M. 1579
 Finn, C. P. 1659
 Finzi, C., 1526
 Fischer, A. J. 1610
 Fischer, H., 1520, 1549, 1567
 Fischer, O. 1572
 Fischer, R., 1632
 Fisher, N. R., 1641
 Fishwick, C. L., 1678
 Fitch, A. A., 1435

 Fitterer, C. R. 1476
 Flaherty, G. P. 1463
 Flansby, M. 1624
 Fleischmann W. 1564
 Florentz M. F. 1604
 *Fletcher W. B. 1642
 Florentin P. 1566
 1583
 Florians L. 1636
 Fodor A. 1543
 Foch's V. G. 454
 Foley S. J. 1548
 Felling A., 4570
 *Fonseca, E. L., 1416
 Fontaine 1594
 Fontès G. 1534 1575
 Fontes J. 1568
 *Forsman, R. A. 1416
 Forrest H. O. 1475
 Forsberg A. 1554
 Forster, A. 1591
 Forster, R. B. 1423
 *Fournet et Lachurée (S. A. R. L.), 1651
 Fouts C. C. 1604
 Fowle, P. D. 1678
 Fowler, O. J., 1617
 Fowler, R. H. 1157
 Fox, C. S. 1164
 Fot, J. T. 1579
 *Fränkel & Viebahn 1416
 François, M., 1159
 Frank, A. H., 1565
 Frank, H. 1582
 Frankel, M. 1447, 1543
 Frankenthal, L., 1543
 Fraps, G. S. 1616
 Frazer, J. C. W., 1432
 Fred, E. B. 1627
 Frederick, R. C., 1635
 Frédéricq, H., 1586
 Friedland, V. D., 1675
 Freeman, M., 1543
 Freeman, S. B., 1634
 Freise, F. W., 1654
 Fretwurst, F., 1573
 Freudenberg, K., 1495, 1496, 1497
 Frey, R. W., 1703
 Frey, S., 1572
 *Frey, W., & Co., 1452
 Friederich, W., 1456
 Friedhelm, E., 1585
 Friedmann, W., 1648
 Friedrich, K., 1493
 Frink, R. L., 1647
 Frisz, J., 1479
 *Fröland, P. E., 1448
 Frölich, A., 1705
 Fuchs, E., 1449

 Fuchs G. 1567
 *Fuchs, O. 1537
 Fuchs, P. 1456
 Fuchs, W. 1656
 Fühner W. 1534
 Futh O. 1544
 Futh R. 1434
 Fukami S. 4564
 Fuzita S. 4270

 Galsano P. 1414
 Gally V. L. 4422
 *Gale M. 1629
 *Gahleit P. J. 1673
 Gale, V. 1531
 *Gardner H. A. 1673
 Gardner H. A. 1654
 1690
 Garnier M. 4555
 Garnier R. 4631
 Garrett O. P. 1594
 Gause A. W. 1469
 1657
 Gavia J. 1543
 Gavilla, J. 1571, 1575
 Gayda T. 1584
 Gehmert Fulness P., 1510
 *Gelsch Helmer Maschinenfabrik, 1415
 Gedy, G. R., 1439
 Gehlen H., 1440
 Gelmann, W., 1157
 Geissler K., 1603
 *Gelatinose Industrie et Vitrode Soc. anon. 1704
 Geldrich, J., 1563
 Gelpi, A. J., Jr., 1594
 Gernant, A., 1463
 Gensud, P., 1550
 *General Technical Co., Ltd., 1668
 George, F., 1572
 Gerke, M., 1654
 Gerecs, A., 1498
 Gerlach, 1619
 Gerold, E., 1474
 Gerthsen, C., 1434
 *Ges. für Linde's Eismaschinen, A. G., 1605
 Gessner, O., 1582
 *Gewerkachalt Eisenhütte Westfalen, 1416
 *Gewerkachalt Mathias Stinner 1538
 Gerehns, R. A., 1436
 Gheorghiu, C. V., 1506
 Ghiron, D., 1445
 Ghosh, S., 1650
 Giacalone, A., 1507

 Gills W. B. 1452
 1674
 *Gibson G. H., 1610
 Gibson T. 1613
 Gibson T. W. 1641
 Giesche F. 1611
 Girard P. 1617
 *Gibbert M. 1415
 *Gillmeister A. 1440
 Gillmeister F. 1639
 *Gill A. H. 1654
 Givett H. W. 1473
 Gibson J. L. 1467
 Givison S. H. 1677
 Giorgi G. 4579
 Girard P. 1410
 Girardot M. 1566
 *Giroud F. P. C., 1616
 Givré M. 1545
 Givandon J. 1655
 Glass J. 1547
 Glass S. P. 1709
 Glasford J. 1621
 Glanville S., 1144
 Glaubitz, 1626
 Glarner A. (Mlle.), 1518
 Glarunov, A. 1412
 Gschütz, L. S., 1513
 Glocker, C., 1410
 Glyde H. S., 1664
 Goebel P., 1551
 *Goetsch, C., 1145
 Gokke, H., 1515
 Golar, H., 1459
 Goldschmidt, S., 1534
 *Goldschmidt, T. A., G., 1537
 Goldschmidt, V. M., 1417
 Goldstein, L., 1454
 *Goodrich, R. J. 1539
 *Gotham, G. 1674
 Gortner, R. A., 1546
 Gordon, J., 1494
 Goto K., 1532
 Gottfried, C., 1460
 Goubeau, J., 1454
 Gould, H. W., 1641
 Grabar, P., 1574
 Graham, H. C., 1417
 Grange, R. H., 1575
 Grant, G. A., 1631
 Grant, R. J., 1641
 Grawe, O. R., 4469
 Gray, A., 1463
 Greaney, F. J., 1622
 Grebl, A., 1654
 Green, E. L., 1426
 Greenhill, A. W., 1602
 *Greer, E. S., 1650
 Grest, K., 1602

- *Grenaudier L, 1671
 Gressig R, 1652
 *Gretcher E F 1650
 Griffith M, 1678
 Grigaut A, 1566, 1575
 Grishspan, L B 1619
 *Grissbaum, L. D., 1667
 *Groeck, H 1610
 Groff J T 1540
 Grossfeld J 1577
 Grassman M A, 1480
 Grosskorsky E, 1649
 Grotman W 1441
 Grounds A 1656
 Grove Palmer, P., 1678
 Grubb W J, 1511
 Grube G 1121, 1432
 Gruber, C. M., 1591
 Gruener, H W, 1433
 Grönwald M, 1632
 Gruner, J W, 1465
 Gruner, P 1433
 Guarnieri, P 1454
 Guawh, J R, 1458
 *Gubelmann, I, 1539
 Guéhen, G, 1436
 Guenther, A, 1437
 Guenther E. S., 1631
 Günther, J, 1625
 Guba, P C, 1503
 Guillemet, R., 1459
 Gufflet, L., 1474, 1476
 Guilleminot, 1450
 Gurvitch, A, 1544
 Guy, P 1476
 Gwoods J 1473

 Haag J R 1561
 Haav K 1637
 Habu H 1599
 Haddon, E., 1690, 1700
 Häfner, I, 1545
 *Hägglund, E, 1670
 *Häble, H, 1696
 Händer, H, 1581
 Hagenbach A 1439
 Hager, G, 1612
 Hahn K, 1678
 Hahn, Q., 1471
 Haad R., 1628
 Hale, F E, 1607
 Hale O A 1620
 Halford J O 1429
 *Halquand, G 1673
 Hall A J 1677
 Hall G E, 1553 1560
 *Hamburgische Gaswerke G m b H 1643
 Hamilton, B H, 1473
 Hamy, R, 1695
 Hanak, A, 1601
 *Handelsges, W., 1415
 Hanemann, H, 1474
 Hanko, M T, 1537
 Hanna, W F, 1622
 Hansa H C, 1593
 Hanson, S, 1540
 Hardman A. P., 1705
 *Hardy, H, 1640
 Harnamowitz H, 1471
 Hartman, P A, 1534
 Harris E E, 1469
 Harris P C, 1441
 Harris G H, 1552, 1676
 *Harris, J P, 1646
 *Harriss, J E G, 1633
 Harris R H, 1594
 Harrop, G A., Jr., 1582
 Hart L. P., 1691
 Hartford P D, 1603
 *Hartmaen H, 1448
 *Harvey, P W, 1625
 Hasama B, 1551
 Haslam, J H 1653
 Haspel, O, 1493 1500
 Hatcher M P, 1607
 Hattoel, S, 1525 1575
 Hettigens I, 1533
 Hlaught J W, 1458
 Hleuonika, P, 1542, 1570
 *Häuser, M, 1651
 *Häusmeister, P, 1346
 Haussmann H W 1634
 Haven G B 1678
 *Haver E 1632
 Hayashihara H 1149
 Hayes W (Miss), 1494
 Hazard R., 1555
 Hecker, T, 1454
 Hedges B S, 1473
 Heide, P, 1470
 Helm de Balsec, F., 1703
 Heim de Balsec, H., 1703
 Henemann, A, 1659
 Henemann A C, 1689
 *Henrich, E, 1625
 Hens, C., Jr, 1479
 Held S, 1414
 *Helt, H 1416
 Heller G, 1516, 1522
 Hellalrom H, 1546
 Hellwig C A 1673
 *de Hen H 1672
 Henderson C N, 1609
 Henley A T, 1629
 Henriqueta V 1580
 Henry, T A, 1512
 Hensel, W G, 1491
 Hepburn, G A, 1621
 Hepner, B, 1597
 Heptostall, W G, 1657
 Herold J, 1632
 Herck, H T, 1554
 Herrmann, K., 1470
 *Herrmann, W O, 1537
 Herwiel W H, 1419
 Herry, C. H., Jr., 1476
 Hertsch, P A, 1650
 Hertsch, W, 1551
 Hertfeld, R, 1590
 Hess, P C, 1183
 *Hess, K, 1603
 *Hessen, R., 1643
 v Hesser, F, 1478
 Heubner, W, 1553
 *Heuck C., 1645
 Heuckekien, H, 1609, 1610
 v Hevesy, G, 1137, 1471
 Hey, M H, 1461, 1162
 Heymans C, 1571
 Heymann J P, 1571
 *Heyna H, 1683
 Heyrovsky, J, 1441
 Hicks, C. G, 1599
 Hicks, W M, 1433
 Hida T, 1531
 Hiebert P, 1100
 Hildebrand, J H 1427
 Hill, A V, 1546
 Hill E, 1568 1569
 Hill H J, 1653
 Hill, H. S., 1608
 Hile, J, 1421
 *Hillebrand, H, 1660
 Hlodie, E, 1579
 Hlossberg, G, 1417
 Hlons, G, 1434
 *Hirsch, Kupfer und Messingwerke A.-G., 1431
 Hurt, A, 1443
 Hutschen C S 1453
 *Hudson, T E, 1604
 Hochrein M 1581
 *Hodges A B, 1610
 Högl, F 1581
 *Höller, K F 1645
 *Hofka, E, 1655
 Hoffmann, P, 1564 1639
 Hoffmann, G, 1695
 Hoffmann E, 1507
 Hoffmann H E, 1690
 Holck, H G O, 1562
 Holden, H F, 1543
 Hollen, J H, 1697
 Hollands H W., 1655
 *Hollenberg R W., 1610
 Holfers, M L, 1694
 Hollett, G T, 1603
 Holmes, B, 1579
 Holmes H N, 1433
 Holthaus, C., 1479
 Holts, P, 1572
 Holmann, E, 1663
 Homme, S, 1593
 Hooncamp P, 1625
 Honda, K, 1477
 *Hoonay, C., 1661
 Hopkins, G R, 1657
 Hopkins H C, 1667
 Hopkins, R H, 1624, 1629
 Hoppe W, 1471
 *Hori S, 1646
 Horn, L. P, 1603
 Horn, O, 1669
 Horn, V, 1558
 Horne W F, 1457
 Horstern, H, 1545 1551
 Houdama T, 1535
 Hotta S, 1565
 Houdremont, E, 1479
 Hough C T 1604
 Housay, D A, 1591
 Hovorka, V, 1452
 Howard J W, 1543
 Howard W V 1469
 *Howell, E T, 1539
 Hoyer, H, 1533
 Hoyer J C., 1561
 Hoyt C. S, 1117
 Haus, L. W, 1501
 *Hubbard, A S, 1447
 Hubert, E R, 1669
 Hubber, W G, 1472
 Hacker G J, 1549
 Häbner, K, 1573
 *Huile, goudrons et dérivés 1663
 Huhn, C. D, 1464
 Hulpreu, H R, 1504
 Hummer, O, 1629
 Humphreys & Glasgow, Ltd., 1661, 1662
 *Hunneman, R. D., 1666
 Hunter, G W, 1431
 Huppert, O, 1473
 Huse, G, 1606
 Husfeldt E, 1540
 Hutscheureuter, R, 1623
 Hykes O V, 1593

- 18nger H., 1633
 *J G Farbenindustrie, 1416, 1431, 1459
 1481, 1487, 1536
 1537, 1538, 1579
 1625, 1626, 1639
 1610, 1642, 1644
 1645, 1616, 1660
 1661, 1666, 1670
 1671, 1672, 1680
 1681, 1682, 1683,
 1684, 1685, 1686
 1687, 1691, 1692,
 1705, 1706
 Ikuta, H., 1695
 Ileva Ratner, R. M., 1629
 Ilge, W., 1420
 Ilung, T. T., 1599
 *Imperial Chemical Industries, 1414, 1538,
 1539, 1612, 1661,
 1662, 1673, 1681,
 1683, 1706
 Inaba, K., 1424
 Inaika, T., 1561, 1583,
 Inahinette, I. H., 1631
 Inneson, S., 1533
 *Iredell, C. V., 1480
 Irosberger, A., 1508
 Ishell, H. S., 1494
 Iskyul, V., 1468
 Isobe, Y., 1477
 Itano, A., 1413
 Ivanov, S., 1704
 Ivkovice, H., 1607,
 Iwase, I., 1425
 Izard, J., 1601
 Izumi, J., 1573

 Jackson, C. V., 1418
 Jacol, K. D., 1619
 Jacol, J., 1582
 Jacobs, V., 1517,
 Jacobs, F., 1705
 Jacobs, M. B., 1428
 *Jacobsen, N. A., 1673,
 Jacobsohn, K., 1420
 Jacobson, C. A., 1438,
 *Jacobson, J., 1640,
 Jahiel, R., 1574
 Jahn, D., 1564
 Jekhola, S., 1582
 *Jakubowitz, M. B., 1614
 Jambuserwala, G. B., 1518
 Jamson, E., 1546
 Jamet, A., 1703
 Jamson, W. A., 1635,
 Jander, G., 1609
 Jannay, J. C., 1635
 *Jansen, H. C., 1630
 Jantach, G., 1451
 Jarrier, F., 1113
 Jarsch, H., 1510
 Jansen, S., 1419
 Joubert, G. P., 1445,
 *Jaulat, G. P., 1633
 Jovet, E., 1601
 Jambor, M., 1561
 *Jeffrey, J., 1631
 Jellinek, K., 1431
 Jenke, M., 1504
 *Jenkins, C. H. M., 1482
 Joanki, N., 1517
 Johanson, A., 1473
 John, H., 1539
 John, C. K., 1597
 Johnson, A. P., 1567
 Johnson, A. H., 1598
 Johnston, R. L., 1376
 Jones, R., 1175,
 Jones, G. H. G., 1614
 *Jones, J. M., 1683
 *Jones, J. T., 1667
 *Jones, P. C., 1706
 Jones, S. M., 1705,
 Jones, T. R., 1613
 Jones, W. J., 1485,
 Jones, T. V., 1410
 Jonston, J. D., 1490
 Joseph, T. L., 1441
 Josephson, K., 1511,
 1534
 Jollyn, M. A., 1709
 *Jourdan, F., 1641
 Jovet, Laverne, F., 1351
 Jüttemann, R., 1571
 Jung, H., 1161
 Jung, L., 1586
 *Junghans, S., 1491
 *Juonen, M., 1671
 Justin - Néangon, L., 1587
 Kämmerer, H., 1572
 Kaganov, I. N., 1709
 *Kahle, H., 1605
 Kahn, R. H., 1582
 Kailen, A., 1508,
 Kaiser, E., 1471
 Kakeuta, K., 1600
 Kakutani, S., 1635,
 *Kali - Forschungs - Anstalt, G. m. b. H., 1644
 *Kalle & Co., A. G., 1452, 1672
 Kallung, B., 1478
 Kalman, E., 1637,
 Karmann, P., 1616
 *Kammerl, H., 1670
 Kanao, S., 1636
 Kanla, J. E. A., 1667,
 Kappen, H., 1619
 Kapustinski, A. P., 1430
 Kargin, V. A., 1423
 Karmashov, V., 1612
 Karmann, H. J., 1620
 Karper, C., 1568
 Karmark, K. A., 1636
 Kartashov, A. K., 1702
 Kata, T., 1464
 Kavokin, F. N., 1619
 Kawamata, J., 1678
 Kehler, L. F., 1567
 Keep, F. T., 1166,
 1467
 Keel, A. H., 1601
 Keh, R. W., 1607
 Kenyon, A. P., 1411
 Kepler, R. J., 1433
 Kermack, W. O., 1505
 *Kessner, M., 1415,
 Kermode, D., 1678
 Kern, F. P., 1443
 *Kere, H. E., 1416
 Kleff, F., 1473
 de Kiewiet, T. F., 1503, 1507
 Kiki, S., 1702
 Kimball, W. S., 1419
 Kinnus, K., 1573
 *King, A. T., 1681,
 1684
 King, C. V., 1128
 King, Y. J., 1518, 1535,
 1560
 King, H. S., 1415
 Kinoshita, K., 1426,
 1552
 *Kinsley, C., 1491
 Kipping, P. B., 1481
 Kiro, E. R., 1427
 Kirsch, R., 1586
 Kiro, H., 1675
 Kitawato, Z., 1532
 Klason, P., 1665
 Klatt, P., 1632
 Klebeberger, 1619
 Klein, M., 1586
 Kleinschmidt, H., 1573,
 *Kleissner, W., 1663
 Klenk, E., 1541
 Kling, K., 1636
 Koppel, J., 1516
 Klopp, W. J., 1451
 Klüber, H. v., 1431
 Klumb, H., 1434
 Knöfel, J., 1428
 Knott, J. E., 1614
 Knuchel, H., 1653
 Kobayashi, T., 1477,
 Kobel, M., 1545, 1553,
 1527,
 Kobitsu, K., 1700
 Koch, H., 1620
 *Kohle K., 1684
 Köhler, F., 1461, 1468
 Köhler, P., 1547
 Köhler, K., 1640
 Köhring, V., 1543
 Koenig, E. W., 1617
 Köfer, A., 1509
 Köfer, L., 1509, 1670
 Kohlschütter, A., 1441
 Kohren, F. P., 1599,
 1600
 Kullath, W., 1555
 Kondo, J., 1668
 Koss, V., 1701
 Korala, A. I., 1658
 Korol'ski, A., 1626
 Korkowski, W., 1571
 Kosteritz, H., 1582
 Kötting, P. R., 1479
 Kotto, S. A., 1707
 Kostyl, I., 1548
 Kounts, W. H., 1701
 Kouritsky, H., 1587
 Krawenkaya, T., 1668
 Krawtshchikov, H. I., 1701, 1707
 Krause, H. F., 1471
 Krause, W. E., 1557
 *Krause, R., 1415
 *Kraus, Monokotto Vereinigte Industrieanlagen A. G., 1630
 *Kraus, H., 1642
 Kraut, H., 1581
 Kraybill, H. R., 1551
 Kraut, F., 1650
 *Kraupl, A., 1603
 Kreulen, D. J. W., 1417, 1656
 Kringstad, H., 1500
 Krishnamurti, K., 1126
 Kroeber, L., 1631
 Kroll, S., 1632
 Krüger, 1660
 Kruisbeek, C. I., 1597
 Kryl, P., 1701
 Kryszykiewicz, J., 1657
 Kühl, H., 1595,
 *Kuenemann, T., 1448,
 Korschner, K., 1601
 Korschner, O., H., 1632
 Kuelmann, B., 1572
 Kuhn, W., 1496
 Kukharenko, I. A., 1700
 Kuhlmann, G., 1632
 Kuhn, R., 1513
 Kuk, S., 1487, 1543
 Kulp, M., 1439
 Kumagai, T., 1564
 Kunitz, W., 1461
 Kunsman, C. H., 1616
 Kuntz, R., 1595

- Nitescu, I. I., 1553,
 1593
 Niven L. A., 1622
 *Noar R. J., 1706
 Noll, L. O., 1559
 Noll W., 1662
 Nolte, O., 1519 1620
 Nord P. P., 1546
 Normann W., 1693
 Norris, F. M., 1616
 *Norton Co., 1631
 Notbhaft, J., 1421
 Notman A., 1641
 *Novak, I. J., 1646
 *Novaseta A. G. Ar
 bon 1672
 Nogent R. L., 1540

 Oberdorfer, 1616
 Ohreaga, A., 1506
 *d Obry, V. W., 1673
 *Ocean Steam Ship
 Co., Ltd., 1416
 Oesper, R., 1417
 Oettinger J., 1703
 Ozata, A., 1559
 Ochiai, S., 1560
 Ohl, P., 1603
 *Ohlendorf Dolge, J.
 S., 1534.
 Okawa K., 1564,
 1552
 Okano K., 1594
 Oon, M., 1510
 *Optisch-mechanische
 Fabrik Steindorf &
 Co., 1545.
 *Oriandi, O., 1445
 Orion, E. Jr., 1648.
 Osborn T. W. B., 1414
 Osborn, W. A., 1543
 Ostendorf, C., 1545
 Ostinell, F., 1524
 Ostrogovich, A., 1531
 *Ostromislensky, I.,
 1539
 Otsuka I., 1589
 *Otto, C., & Co., G. m.
 b. H. 1662
 Overman, O. R., 1528
 Owen F. V., 1620
 Owen W. L., 1417
 Oxford A. E., 1515

 Pace E., 1443
 Page I. H., 1541,
 1576
 Paget, H., 1512
 Paget M., 1571, 1637
 Pagnell A., 1604
 Paic M., 1431
 Paleček E., 1701
 Palkin A. P., 1429
 Papadanco, O., 1643
 Paneth P., 1418
 Pannockock, A., 1441
 Pantaselli, E., 1694
 Pantelejev, I., 1694
 Pantia, C. P. A., 1523
 Paay, J., 1544
 *Pardoen O., 1633
 Parisi F., 1609
 Park C. P., Jr., 1463
 Parker, A., 1638.
 *Parker, L. D., 1633
 Parkinson, S. T., 1602
 Parr, S. W., 1657
 Parravano, N., 1432,
 1616
 Parrish P., 1660
 *Parsons, C. A., 1415.
 Parsons, H. T., 1559
 Paschal, P., 1492
 Paschuk K., 1564
 Pasqualino O., 1519
 Pasturini, 1579
 Patel, J., 1535
 Patnode, W. I., 1454
 Patterson, T. C., 1635
 Patwardhan, G. B.,
 1594
 Patwardhan, V. G.,
 1615.
 *Paul, A. Jr., 1651
 Paulsg L., 1420
 *Paulsen W. R., 1450
 Payman, W., 1575
 *Péchin, L. C. F.,
 1597
 Pedersen K. J., 1431
 Pedersen, C. S., 1549
 Penfold, A. R., 1631
 *Pepodest Co., The,
 1640
 Perkins A. G., 1516
 *Perks, E., 1603
 Peters, H. S., 1621
 Peters, R. A., 1547
 Peterson W. H., 1637
 *Petit, C., 1662
 Petit, P., 1628
 Petit, V., 1610
 Petow, H., 1582.
 Petrov, G., 1691, 1694
 Peyer, W., 1633
 Pianbauer, J., 1658
 1659
 Pfeiffer, H., 1547
 Pfeiffer, P., 1510 1533
 von Philpsborn H.,
 1460.
 *Phosphate Recovery
 Corp., 1643
 Phragmfa G., 1431
 *Physiologisch-chem.
 isches Laboratorium
 Hugo Rosenberg
 1649
 Mattli, L., 1419
 Pichagina, A., 1694
 *Pickard, H., 1653
 Picet, A., 1424
 Pieters H. A. J., 1555
 von Plad, S., 1663
 Pincus, H., 1667
 Pincusman L., 1444
 Piper, G. R., 1611
 *van der Plas J.,
 1625
 Planner, R., 1443
 Plant S. G. P., 1522
 *Platt, H., 1656
 Plarek, E., 1504
 *Plomshier, A., 1645
 Plarher, W., 1544
 Plescherff G., 1510
 Polak, R., 1635
 Polcard, A., 1593
 Pollak, H., 1593
 Pollak, J., 1505
 *Pollitzer, P., 1605
 Polonovski M., 1565,
 1575
 Polynov, R., 1614
 Popp, A., 1475
 Poppo, G., 1459 1491
 Popberger, P., 1625
 Pope, W. T., 1600.
 Popesco, M., 1563
 Popescu, H., 1580
 Popoviciu, G., 1590
 Popp, W., 1622
 Popper, H., 1545, 1575
 Popper, L., 1592
 Porin J., 1585
 *Porter E. A., 1414
 Post H., 1433.
 *Posental A., 1415
 Potter F. M., 1534
 Potter G. F., 1531
 *de Potter, L., 1667
 *Potta, C., 1611
 *Pourelais, R., 1653
 Pouret A., 1676
 Powell, C. K., 1599
 *Powell, H. M., 1635
 *Power Gas Corpora
 tion, Ltd., 1662
 Prakash, S., 1580
 Pratt, T. W., 1554
 Prescott J. A., 1611
 Prescott S. C., 1610
 Pretscher P., 1449
 Pretschner F., 1450
 Priddy, E., 1574
 Price D. J., 1675
 Priestler R., 1632
 Prikl, S., 1582
 Pringheim, H., 1494,
 1495
 *Products Roche, Soc.
 anon., 1537
 Proescher, P., 1575
 Prosser White, R.,
 1604
 Protopopescu L.,
 1511
 Publow, H. P., 1475
 Pucher, G. W., 1553
 Pucherna, J., 1697
 Puiggari, H., 1634
 Purdy, C., 1567
 Puri, A. N., 1614
 Pyman, F. L., 1530

 Rabald E., 1414
 Rabald, J., 1549
 Rabinovich, A. J.,
 1423
 Rabinowitch I. M.,
 1571
 *Rabjohn L. E., 1610
 Radarovich, R., 1573
 Radcliffe C. D., 1516,
 1631
 Radima, W., 1567
 Rader, M. G., 1445.
 Rajha, C. F., 1589
 Ramin, R., 1584
 *Rak A., 1701
 Ramage G. R., 1620
 *Rambush N. E.,
 1652
 Rame A., 1641
 Ramnathan, 1575
 Randall, M., 1682
 Randow, L., 1559,
 1567
 *Rankin, W. A., 1418
 Rapkin L., 1594
 Rasmussen, E., 1439
 *Rasquin, P., 1667,
 Rathery, P., 1587
 Rathner, E. I., 1618
 Rautenbusch W.
 1664
 Ray, J. C., 1462
 Ray, S., 1419
 *Ray, W. R., 1416
 Rayleigh, L., 1438.
 Raymond A. L., 1493
 Raymond - Hament,
 1590
 Reynolds, E. L., 1616
 Rayner, J. W. R., 1456.
 Re U., 1664
 Rea H. E., 1553
 Read J., 1511
 Rebmann O., 1549.
 Reclus, H., 1667
 Redmond, J. C., 1456
 Redlob, E., 1586
 1583
 *Reedy A. J., 1416.
 Rees, W. J., 1645.
 Rehder, W., 1512

- Reiche, P. D. 1573
Reid C. 1563
*Regnier M. 1416.
Reilly, J., 1494 1499,
1513 1523
Rein C. 1473
Reiss J. 1519
Reiss M. 1570 1590
Reiss, P. 1566 1583
Reith J. F. 1580
Remond A., 1574
Renaud M. 1575
Rendon, Q. D., 1700
Reo M., 1449
Resnikova, S. B., 1704
Reutershild J. A.,
1457
*Revos V. 1451
de Rey Paulhade, J.,
1540
*Rewald, B., 1636,
1704
Ricard, P., 1533
*Riccardi, G. S., 1656
Ries, O., 1678
Richard, L., 1574
Richards, F. J., 1609
Richardson, C. H.,
1622, 1624
*Richardson, W. H.,
1674
*Riehr, F., 1653
*Richter, G. A., 1673
Richter, G. H., 1590
Riebel, R., 1705
*Riedel, J. D., *E. de
Haën, A. G., 1640
Rieger, F., 1660
*Rieger, W. F., 1415
Rienäcker, G., 1549
Rigó, L., 1633
*Riley, R., 1443
Rindt, E., 1589
*Ring Ges. chem. Un-
tersuchungen m. b.
H., 1415
Ripley, L. B., 1621
Ritter, J., 1668
Roark, R. C., 1621,
1670
*Roark, R. C., 1626
Roberts, H. C., 1677,
Roberts, H. S., 1459
Roberts, J. G., 1441,
Roberts, J. K., 1466
*Roberts, J. W., 1626
Robertson, A., 1522
Robertson, J. D., 1570
Robiette, A. G., 1442
Robin, J., 1501
Robinson G. W., 1612
Roche, A., 1580
Roche, J., 1555
Roemans, M., 1586
Rodenkirchen, J., 1535
*Rodgers, S. T., 1603
Rösner, K., 1442
Roethlis, H. F., 1478.
*Roffey, P., 1662
Rogers B. A., 1479
*Rogers P. M., 1666
Roug A., 1553
Rolle W. A., 1622
Rollet A. P., 1445
Roman W., 1584
Romanoff, A. L., 1565
Romer, J. R., 1472
Roncato, A., 1569
Rose, G. R., 1473
Rosen, B., 1690
Rosenarten C., 1573
*Rosenhau, W., 1482
Rosenthal, D., 1476
Rosenthal, F., 1572
Roseenthaler, L., 1632
Rosen, P., 1656
Rosselund, S., 1441
Rosenbeck, H., 1572
Rösner, G. A., 1431
Roth, H., 1562
Rother, H., 1609
Rothschild, F., 1564
Roukheiman, L., 1583
Roule, L., 1610
*Roure, M., 1621
*Rowell, S. W., 1449
Rowntree, J. I., 1556
Rubens E., 1674
de Rudder, P., 1430
Rude J., 1659
Rudolfs, W., 1609,
1610
Rue, H. F., 1663
Rüfer, R., 1595
*Ruhemann, F., 1633
Ruechjelm, D., 1554
Rupp, E., 1433.
*Russ, E. F., 1418
Russocki M., 1503
Rutherford P., 1437
Rutler, J. E., Jr.,
1426
Ryan, E. M., 1526
Ryan, R. J., 1593
Rydbom, M., 1546
Rysselberghe, P. J.,
1443
Saakov, S., 1631.
Sabatschka, T., 1638
Sach, J. S., 1638
Sachse, E., 1673
*Sägebarth, B., 1415
Sagara, J.-I., 1545
Sager, B., 1582
Sabyun, M., 1543
Sailard, E., 1701
St John, J. L., 1426
Sainton P., 1575
Sakai, H., 1704
*Saklatwalla R. D.,
1442
*Sala C. J., 1645
Salceon C., 1415
Salstrom, E. J., 1427
Salzmann C., 1475
Sampson A. W., 1551
*Samuel A., 1692
Sanborn N. H., 1600
Sand H. J., 1457
Sandels M. R., 1557
Sänders, K., 1697
1699
Santesson C. G., 1582
Saradjevich, F., 1581
Sarkar, P. B., 1455
Sartini, L., 1526
Sasaki, T., 1559
Saso, T., 1578
Sato, M., 1694, 1695
*Saxrefabrik
Schweizerthal,
1645
Scarpis O., 1446
Schacht, P., 1614
Schachtel, P.,
1461
Schaefer, W., 1664
Schäfer, A., 1570.
Schaphorst, W. F.,
1604
von Schiele, C., 1702
Schemansky, P., 1564
Scheyer, F., 1500
*Scherer-Kahlbaum,
A.-G., 1534, 1610
Schreier, M., 1627,
Sebeunert, A., 1556,
1558, 1562
Schiedt, E., 1432
Schiff, A., 1553
Schiff, P., 1534
*Schiffner, H. J., 1481
Schiffman, E., 1632
Schikorr, G., 1609
Schill, I., 1489
*Schinke, F., 1480
*Schirmacher, K., 1653
Schlemmer, O., 1592
*Schless, S., 1645
Schloss, Wm., 1593
Schlossberg, T., 1558
Schmidt, L., 1473
*Schmid, M., 1680
*Schmid, R., 1691
Schmidt, P., 1422,
1572
*Schmidt, H., 1645
Schmidt, H., 1670
Schmidt, H. H., 1449,
1450
Schmidt, J., 1516
*Schmidt M. (Miss),
1651
Schneider, P., 1555
Schneider P., 1516
Schneiderhöhn H.,
1465
Schoen J. (Mlle), 1519
Schoen, M. J., 1676
v. Schoenbeck O., 1541
Schottky W., 1431
Schour M., 1573
*Schroder G., 1416
Schultes W., 1655
Schulz K. G., 1627
Schulze, R., 1656
*Schur M. O., 1673
Schurz, P., 1664
Schuster, P., 1655
Schuster, G., 1637
Schut, W., 1480
*Schwartz, G. L., 1671
Schwartz, G. M., 1463
Schwarz, P., 1665
Schwarzenbach G.,
1425
Schweitzer, A., 1594
Schweitzer, H., 1597,
Schwyzer, J., 1632
*Scorra, C., 1416
Scott, E. B., 1611
*Scott, J. W., 1662
Scott, W. M., 1676
Scott, W. W., 1631,
Severnor, J. B., 1467
Seattle, A. D., 1650
Second Commission
Internat. Soc. Soil
Sci., 1614
Seckles L., 1576
Seidell, A., 1561
Seiler, W., 1627
Seifriz, W., 1540
Selkin, R., 1658
Selye, R., 1589
Semachon, L., 1628
*Semon, W. L., 1705
Senécal, L. H., 1639
Sendo, M., 1668
Sendroy, J. I., 1567
*Senen, C., 1419
Sepilli, A., 1561
Sergeev, M., 1605
Seto, I., 1605
*Seitzler, H. R., 1666
*Severini, O., 1416
Sevilla, M., 1691
Sevin, E., 1434
Sewell, W. E., 1565
Seyewetz, A., 1449,
1450
Seyfarth, M., 1508
Shakhno, A. P., 1659
Shappell, M. D., 1420

- Sharkor V I 1669
 Sharp F F 1509
 Shaw W N 1617
 *Sheafor D L 1673
 Sheard C 1567
 Shedd O M 1613
 Sheftel A G 1549
 Shephard H H 1622
 Shephard F C 1632
 Shepherd M 1455
 Sheppard S E 1419
 Sheridan L M 1472
 Shetman P 1623
 Sherman H C 1557
 Shibata K 1528
 Shimizu J T 1490
 Shinoda Y 1669
 Shioneka K 1636
 Shison M 1419
 Shively W L 1659
 *Shoemaker R J, 1452
 Short M N 1166
 Short, W P, 1516 1631
 Shukla S N, 1146
 Shuya S, 1535
 Sushar G, 1562
 Siegfried K., 1635
 *Siemens & Halske A. G., 1147, 1419
 *Siemens - Schuckert werke A G., 1605
 Sievers A F, 1633
 *Silbermann E., 1444
 *Silberrad O, 1671
 Simon D, 1562
 Simon C 1641
 Simonnet H., 1560 1575
 *Simonsen J L, 1631
 Simpson B T, 1580
 Simpson E S 1460
 Simski A., 1621
 Singer K. 1474
 Sirtaki M 1702
 Sudo, K 1669
 Suley, J P 1677
 Sivolobov A. V., 1658
 Systeina D, 1578
 Shalla, N 1454
 Skinner H W B., 1435
 Slater L. 1659
 Slater, R. H., 1528
 Slater S M 1643
 Stepan J 1435
 Smith C. R., 1622
 *Smith Drum & Co., 1656
 Smith J 1417
 Smith J F 1505
 Smith L 1485
 Smith M A. 1624
 Smith M I, 1561
 *Smith, R D. 1650
 Smith R H, 1624
 Smith S, 1533
 Smith S W, 1472
 Smith W C. 1606
 *Smoot Engineering Corp 1415
 *Smorodintzev, A., 1549
 Smapp O I 1625
 Snell A. M 1577
 Sohyama N 1631
 *Société lair liquide (Soc anon pour l'étude et l'exploitation des procédés Georges Clavel) 1603
 *Société anon des anciens établissements Skola à Posen 1451
 *Soc anon, la Carboneite 1643
 *Soc. anon des foyers automatiques, 1413
 *Soc anon pour l'ind chim à Hâle, 1449 1670 1674 1691
 *Soc. anon des mapp factures des glaces et produits chim de Saint - Gobain Chauny & Crey, 1650
 *Soc anon, la nouvelle montagne, 1690.
 *Soc. chim de la Grande Paroisse (Acide et produits chim) 1604 1643
 *Soc. d'électrochimie, d'électromé talurgie et des sources électriques d'Ugine 1451
 *Soc d'études pour la fabrication et l'emploi des engrais chimiques, 1625 1643
 *Soc d'études scientifiques 1626
 *Soc d'études scientifiques et industrielles 1449 1697
 Soc Fleury et Lavanant, 1704
 *Soc générale de fours à coke (Systèmes Lecroq) 1644
 *Soc Henry Simon, Ltd 1414
 *Soc industrielle de Mouy, 1671
 *Soc Italiana Fardelli, 1705
 *Soc. Kodak Pathé (Soc anon. française), 1671
 *Soc Lévy, Samuel & Lévy, 1692
 *Soc de recherches et d'exploitations pétrolières 1665
 *Soc des unnes chim Iquaa R h o n s . Poulenc 1562
 Sulzer K., 1423
 Sulas F L., 1566
 Sureau C. H., 1125.
 Soudet J, 1702
 Soula, C., 1567, 1568.
 Southcombe, J F., 1665.
 *South Metropolitan Gas Co, 1653
 Spangenberg K, 1451
 Spaulding C. H., 1607
 Speckman J C., 1444
 Spencer, B, 1460
 Speisler, O., 1698 1699
 *Sperr, F W., Jr, 1661
 Speter, M., 1417
 *Sprague, A. E., 1650
 Spöel 1151
 *Standinger, H., 1679
 Stashikov G, 1472
 Staemmler, C., 1659
 Stahl W, 1435
 Stager 1626, 1627
 Stakhoraka, V M., 1640
 Stammers A. D, 1411
 Stampe G, 1413
 *Standard Oil Development Co, 1626
 Stanek V, 1701
 *Stanner, H., 1653
 Stanley H M, 1664
 Stanton T E, 1665
 Stareck J F, 1423
 Stark C., 1690
 Starkiewicz J, 1440
 Stanak, A., 1633
 Stauder F, 1632
 Steffen, W 1609.
 Steff J, 1567
 Steger A, 1693
 Steinberg P, 1564
 Steinbrunn, G., 1498
 Steiner, W, 1429
 Stemmetz H, 1421
 *Stellawerk, A G vorm Wtlach & Co, 1415.
 *Stephen, W L N M., 1631
 Stephen, H., 1501, 1505 1507, 1511
 Sterner Raiser, L., 1450
 Stillbacher, A., 1674
 *Stettiner Chamotte-Fabrik A.-G. vorm. Dindler 1662
 Stetsenko I., 1694
 Stewang W, 1439
 *Stevens, J H., 1651
 Stewart, C. P., 1555.
 Scheeman A, 1540
 *Stich E G., 1605
 *Stid C., 1662
 Stillwell F L., 1464
 *Stimson P E, 1615
 Stobbe H., 1441
 Stock A, 1456.
 Staltesberg - Bergius, M., 1604.
 Stoltzenberg H., 1601
 Slowe, V M., 1177
 Siroli & Cie., 1696
 *Straub, F., 1640 1651
 Strauss, S. D., 1641
 Streck H., 1133
 Stevens, J L., 1661
 Stroganova Z I., 1615
 Sturdivant, J H., 1420
 Sturivant, H B., 1677
 Sucher L., 1515
 Suis O 1520
 Sagai K 1564
 Sugden S., 1185
 Sugi K., 1470
 Sullivan J T, 1551
 Sumi M, 1534
 Sumami S., 1573.
 *Supren, C, 1662
 *Sutcliffe T R, 1628
 Suzuki K., 1555
 Svodberg T., 1546
 Sveshnikova V I., 1704
 Swallow S. C. H., 1645
 Swan J N 1417,
 Swann, S., Jr, 1456.
 Sward, G G, 1680
 *Sweetland, E J., 1667
 Swietoslowski, W, 1659
 Szakili, A, 1563
 Szancor, H 1632
 Szegö L., 1524
 Tadokoro, T. 1554.
 Tafel K., 1693.

- Taft A E 1510
 Taft R 1423
 Tainton F C 1442
 Tait P G 1742
 Takata H 1535
 Takeda K 1573
 Takel S 1510
 Taku A 1548
 Talbot H F 1452
 Tammer 1440
 Tamya, H 1531
 1549 1550
 Tammann G 1426
 Tanaka K 1531 1549
 Tanaka M, 1659
 Tanida, J, 1428
 Tanret, G, 1552 1560
 Tanser, P 1672
 Tapie, J, 1543
 Tasman, A, 1595
 Tatum, A L, 1591
 Taylor, H S, 1422
 1663
 *Teatini, D, 1703
 Temesvary, I, 1633
 Tenenbaum, M,
 1542
 Teorell, T, 1543
 Terauchi, K, 1582
 *Terry, C. M., 1415
 *Tersian H G, 1662
 Tesche, F, 1443
 Tester, A C, 1413
 Tewes, K, 1478
 Theophilus, D R.,
 1698
 *Thermen, A J, 1631
 Thews, P R, 1477
 Thibaud, C, 1677
 Thiere H, 1572, 1574
 Thirlar, P, 1606
 *Thirring, H, 1447
 Thivolle, L, 1558,
 1575
 Thomas C A, 1631
 Thompson, G W,
 1472
 Thompson, M deK.,
 1445
 Thompson M R,
 1639
 Thompson, T G, 1604
 Thoms H, 1633
 Thomson, F, 1466
 Thomson, J R, 1625
 *Thorne, G, 1704
 *Thornton, R H,
 1416
 Thurber, D S, 1530
 Tietze, W, 1447
 Timon David, J, 1592
 Tlukow, D, 1553
 Tobey, E R, 1620
 Tödt, P, 1698, 1699
 Tomaszewski S, 1682
 Tonnet J, 1547
 Tougarinoff B 1458
 Tournel W, 1542
 *Tournel M 1692
 Trautwein K 1629
 Trebley I O 1702
 Treibz W 1512
 Treloar A F 1595
 Treves M 1603
 Trillat, J J 1664
 Trofimova P I 1704
 Trotman S R 1677
 Truttwin H, 1639
 *Turner C 1661
 Turner C W 1565
 Turner W P S 1647
 Turowska I (Silke)
 1548
 *Turrentine, J W 1641
 van Tossenweck, M
 J, 1676
 *Twiss, D P, 1674
 Tyler, T M, 1641
 Tyson J, 1619
 Teutnerstver, M,
 1430
 Uchocka S 1574
 Uhrmann, C. J, 1648
 Ukai, S, 1564
 Ulbrich M 1627
 *Ulrich, L. H, 1633
 *United States Metals
 Refining Co., 1448
 *Universitäts - Insti-
 tut für Physikalische
 Grundlagen der
 Medizin, 1643
 Uraki, Z, 1535
 Urbach, P, 1562
 *Urbach, O, 1696
 Usami S, 1695
 Ueuni K, 1560
 Uytvanck, P V, 1596
 Vácha, J, 1700
 Vacher, H C, 1473
 Valentin, F, 1456
 Van Cerve, R, 1609
 Vanderhinden, L, 1616
 *Van Harpen, N H,
 1705
 *Van Meyer, S L. Jr.,
 1447,
 Van Slyke, D D,
 1567
 Van Weiden, F, 1703
 Varentzav, V I, 1630,
 1639
 Varetton, H, 1549 1664
 Varney, P. L, 1548
 Valitko, J, 1700
 Vatrin, J, 1583
 Vaudin, A. G, 1642
 van Veen, A. G 1547
 Vetch F P, 1701
 Valluz, L, 1544
 Venetkaka, 1642
 Ventre, J 1629
 Venturini G 1526
 Verdenes S 1694
 *Vereinigte Chem
 Werke A-G, 1630
 *Vereinigte Silber
 hammerwerke Het
 sel & Co 1441
 *Vereinigte Stahl
 werke A. G., 1441
 Verweil H J 1435
 Vickery H B 1513
 Vita A, 1654
 Villard P, 1452
 Villaret M 1587
 *Villavaso W J 1667
 Vinogradov A P
 1592
 Vladaco, R, 1562
 Vogt, O, 1479
 Vogt, A, 1414
 Vois, V, 1573
 Volkovich, S I, 1619
 Vondrack, R, 1458
 Vondrák, J 1700
 1701
 Voss H F, 1564
 Votolek F, 1448
 *Vulcan Feuerung A.
 G., 1150
 Vyao, N D, 1599
 Waaser, H, 1617
 *Wager, R H, 1415
 Wagle, F V, 1621
 Wagner, A, 1472
 Wagner, C, 1431
 Wagner P, 1450
 Wagner P, 1629
 Wagner, H, 1695
 Wagner - Janczky, T,
 1513
 *Wake J F, 1651
 Walchschmidt - Leite,
 F, 1543
 Walker, B S, 1635
 Walker, P, 1470
 Walker O J, 1446
 Wallace, E K, 1441
 Walldow, E, 1473
 *Walsh, J F, 1651
 *Walston, J D, 1691
 Walter, C. M, 1657,
 Walton, J H, 1433
 Ward, F. A. B, 1437
 Ware, L. M, 1620
 Warrembourg, H, 1575
 Warren, D B, 1420
 Warren, W W, 1648
 Washburn P W,
 1433
 Wasmer, H, 1591
 Wassermann, J, 1629
 *Wasteris, H J M
 1660
 Wastenson, H, 1636
 Watanabe, S, 1659
 *Waterless Gasholder
 Co Ltd, 1642
 Waters, R B 1522,
 *Watte G 1414
 Wawrzyniak 1651
 Weber, 1610
 Weber, C J 1570
 *Weber, L, 1660
 Wechsel A. te, 1625
 Wedekind F 1502
 Webke, P, 1457
 Weidinger A 1495
 Weil A 1512
 *Weiland H J 1539
 Weimard F P, 1531
 Weismann F, 1427
 Weinstrin, F, 1594,
 West, M, 1564
 Weiss J J, 1699
 Westkamp, H F.,
 1478
 *Weich, H V, 1419
 Welfe, F A, 1534
 Wells, R W, 1621
 Welton, P A, 1618
 *Wendt, G L, 1667
 *Wendt, H D, 1603
 Wentworth, S. W.,
 1551
 *Werfen, A, 1418
 Werthan S, 1648
 *Wewott, W D, 1645
 Westgren, A, 1431,
 1476, 1477
 Westover, G, 1556
 Westphal, W, 1441
 Wewers, H, 1599
 Wheeler, R V, 1654
 *Wheeler, T S, 1642
 *Whetzel, J C, 1697
 White, F G, 1652
 White, P, 1704
 *Whitehead, W, 1635
 Whitman, W G, 1433
 *Wickel, P K, 1673
 Widmann E, 1555
 Wiedenholt, W., 1476
 Wieland, H, 1535
 Wiener, A, 1495
 Wiggam, D R, 1684
 *Wigley, C. G, 1611
 Wijkman, N, 1499
 Wilder, G H, 1661
 *Wilderman, M, 1706
 *Wilke K, 1683
 Wilkins, H, 1485

- Williams D A, 1422
 Williams F J, 1476
 Williams H R, 1417
 Williams R, J, 1561
 Williamson, A T, 1422
 *van der Wijgen P C, 1602
 Wilmer F N, 1447
 *Willshaw H, 1704
 *Willstätter R, 1642
 Wilson A N, 1656
 Wilson B D, 1615
 Wilson I H, 1417
 *Wilson, M, 1611
 Wilson P W, 1627
 *Wilson, R P, 1415 1666
 Winegarden, H M, 1567
 Winks F, 1647
 Winns, H A, 1447
 Winklow, C E A, 1610
 *Winterbauer, C. A, 1646
 *Wintermute, H, 1419
 *Wilsonson Alamed Research Foundation 1630
 Witandt W, 1439
 *Wittenberg L, 1653
 *Wribbing H, 1480
 Wolf, J, 1650
 Woll, L, 1455
 Wolff Pinner, A, 1502
 Wolff H, 1620
 Wood, C E, 1441
 Wood, D R, 1594
 Woodman, H F, 1602
 Woodward H G, 1423
 Woodworth C. W, 1623
 Woog P, 1665
 Word W L, 1654
 Work, R W, 1454
 Wozasek, O, 1576
 Wright C H, 1614
 van Wollen Scholten W, 1669
 Würthn K, 1471
 Wulfschu A, 1514
 Wurm, K, 1441
 *Wurster O H, 1415
 *Wylam B, 1663
 Wynn Williams C. E, 1437
 Xanthakos T S, 1456
 Yakovlev, A G, 1423
 Yamada, A, 1655
 Yamada T, 1669
 Yamamoto R, 1670
 Yanagawa T, 1642
 Yarwera N, 1560
 *Yates W J, 1626
 Yokochi M, 1675
 Yochers M A, 1623
 Young, J B, 1617
 Yurganov V V 1619
 Zachariasen W H, 1461
 Zagami Y, 1559 1561
 *Zabo K, 1693
 Zajac, Z, 1439
 Zakowski J, 1544
 Zakrzewski, K, 1427
 Zale, S, 1597
 Zboray, D, 1633
 Zebkman, I P, 1702
 *Zellstoffabrik Walmhof, 1629 1672
 Zemplén, G, 1626
 Zemyatchenski, P A, 1466
 Zhibn, A I, 1650
 Zhukovskaya, M D, 1659
 Zifferer, R, 1631
 Zikhman . Kedrov, O K, 1617
 *Zook J S, 1417
 Zotta G, 1573
 Zucke F H, 1650
 Zuchmmer, F, 1647
 *Zugmondy, A, 1615
 Zung F, 1540 1564
 Zusmanovich, M V, 1649

CHEMICAL ABSTRACTS

Editor: E. J. Crane, Ohio State University, Columbus, Ohio.

Associate Editors: Elmer Hockett, Helen Gams and Leonard T. Capell

Assistant Editors

M. S. Anderson	H. J. Demarest	P. R. Hawk	L. A. Chase	J. J. Sklines
W. L. Bailett	Allen H. Emery	P. R. Howe	A. M. Patterson	W. R. Stearns
G. H. Barton	W. H. Emery	C. H. Kerr	T. G. Phillips	C. G. Town
Edward Barrow	W. F. Fawcett	H. A. Lepper	A. N. Richards	H. M. Symmes
S. R. Benedict	A. C. Fickner	K. K. Marshall, Jr.	Richard Rimback	P. P. Underhill
P. C. Black	C. G. Link	A. R. Millsten	Alfred Rogers	H. G. Wells
P. L. Browne	C. N. Frier	Harlan N. Miner	C. A. Rosales	C. J. West
C. R. Curren	H. W. Gillett	Chas. B. Morrey	A. H. Rubin	H. T. Wherry
I. N. Dale	R. A. Gwynne	C. R. Munroe	J. P. Schaefer	E. P. Wightman
C. C. Davis	W. T. Hall	W. A. Noyes, Jr.	R. Scherubel	I. C. Witt

CONTENTS

Patents	1	22. The Fermentation Industries	1244
Author Index	1	23. Pharmaceutical Chemistry	1245
1. Apparatus and Plant Equipment	1701	24. Acids, Alkalies, Salts and Sul-	1246
2. General and Physical Chemistry	1713	25. Glass, Clay Products, Refractories	1247
3. Subatomic Phenomena and Radio-		26. Enamels and Other Building Ma-	
chemistry	1722	terials	1248
4. Electrochemistry	1723	27. Fuels, Gas Tar and Coke	1249
5. Photography	1743	28. Petroleum Lubricants, Asphalt	
6. Inorganic Chemistry	1749	and Wood Products	1250
7. Analytical Chemistry	1751	29. Cellulose and Paper	1251
8. Mineralogical and Geological		30. Explosives and Explosions	1252
Chemistry	1752	31. Types and Textile Chemistry	1253
9. Metallurgy and Metallography	1754	32. Paints, Varnishes and Resins	1254
10. Organic Chemistry	1755	33. Fats, Fatty Oils, Waxes and Resins	1255
11. Biological Chemistry	1756	34. Sugar, Starch and Gums	1256
12. Foods	1757	35. Leather and Glues	1257
13. General Industrial Chemistry	1758	36. Rubber and Allied Substances	1258
14. Water, Sewage and Sanitation	1759		
15. Soils, Fertilizers and Agricultural			
Poisons	1761		

PATENTS

A statement giving information as to how to obtain patent specifications and drawings, both United States and foreign, is to be found on page i of *Chemical Abstracts*, 24, No. 8.

Published by THE AMERICAN CHEMICAL SOCIETY, which also publishes *The Journal of the American Chemical Society* and *Industrial and Engineering Chemistry*.

Communications for the editor should be addressed to E. J. Crane, Ohio State University, Columbus, Ohio.

Remittances for annual dues (\$15.00) in the American Chemical Society or for subscriptions to this *Journal* (\$7.50 per year), changes of address, orders for subscriptions or back numbers, claims for missing numbers and all matters intended for the *Proceedings* should be sent to the Secretary, Charles L. Parsons, Mills Bldg., Washington, D. C. Foreign postage is extra, being \$1.50 per year added to the dues of members residing outside the United States, its possessions and the countries of the Pan American Union and \$1.50 per year for *Chemical Abstracts* alone, except that only one half of these amounts is added for journals going to Canada. The countries of the Pan American Union are Argentina, Bolivia, Brazil, Chile, Colombia, Costa Rica, Cuba, Dominican Republic, Ecuador, Guatemala, Mexico, Nicaragua, Paraguay, Peru, Republic of Honduras, Republic of Panama, Salvador, Spain and Uruguay.

The Council has voted that no claims will be allowed for copies of journals lost in the mails unless such claims are received within sixty days of the date of issue, and no claims will be allowed for issues lost as a result of insufficient notice of change of address. "Missing from files" cannot be accepted as the reason for honoring a claim.—Charles L. Parsons, Secretary.

Published by the American Chemical Society, Easton, Pa. Secretary's Office: Mills Bldg., Washington, D. C. Entered as second-class matter at the Post Office, Easton, Pa. Acceptance for mailing at special rate of postage provided for in Section 1103, Act of November 3, 1917, authorized July 13, 1918. Mack Printing Company, Easton, Pa.

AUTHOR INDEX

[An asterisk (*) indicates that the reference is an abstract of a patent]

- Aaronson H A, 1993
 Abelous J F 1902
 Achard C., 1884 1890
 1892
 Achard, G., 1897
 Ackermann, D., 1843
 *Ackermann, F., 2005
 Ackermann, W., 1786
 *Acme Rayon Corp.,
 1994
 Acosta, A. V. 1926
 *Acree, S. F., 1729
 Adam, M., 1894
 Adirkes P 1796
 Adkins H. 1793, 1797,
 1799, 1803.
 Adler W., 2009
 *Adolph G., 1794
 Agaloune, V., 1822
 Agamennone G., 1753.
 Agasse - Lafont, R.,
 1786
 *Akt - Ges. Brown
 Boveri & Cie 1744
 *A. - G. Georg Fes-
 slorffs Salzwirke
 und chem. Fab.,
 1934
 Aiello, G., 1903
 Ainsworth, C., 2009
 *Aktiebolaget Kem-
 laka laterier, 1956
 *Alas, J., 1741
 Albansse A. 1843
 Alberts, W., 1777
 Alcock P H, 1949
 *Alden G. R., 1712
 Alder, K., 1806, 1807
 Aldrich E. W., 1930
 Almans L. V., 1728
 Almann I P., 1760
 Allaire H., 1857
 Allan, J. A., 1771
 Allen, E. V., 1903
 Allen, M. B. 1900
 *Alles, J., 1923.
 *Allgemeine Elektri-
 zitäts-Ges., 1744
 Alling H L 1794
 Allmand A. J. 1736
 Allport N L. 1947
 *Allpress, H. S. 1709
 Allyn W. P. 1867
 Alshberg C. L., 1762
 Altmana A., 1894
 *Altwegg J. 1993
 Amako T H 1868.
 Ambert P., 1862
 Ambrust F., 1707
 Ammon R., 1844
 Amoureux G. 1865
 Anderson A. C., 1756.
 Anderson, A. K., 1858
 Anderson, C. T., 1727
 Anderson E O 1918
 Anderson H H 1902
 *Anderson I B 2002
 Anderson, L. 1758
 Anderson L R. 1867
 Anderson W H 1992
 *Andreas A., 1966
 Andrews, A. C., 1910
 *Andrews C. W. 1976
 Andrews F. 1901
 Andrews, R. C. 1773
 Andrews J., 1909
 *Andrews L., 1712
 *Andsol, A., 2011
 *Anode Rubber Co.,
 Ltd., The, 2021
 Anschütz, W., 1912.
 Anselmer, H., 1913
 Anton R., 1812
 *Apoteker C., 2005
 Arakawa T., 1833
 Archibald J. G., 1939
 Arona H., 1747
 *Arrent, A., 1749
 Ariart C., 1916
 Armann, V., 1754
 Armstrong, E. F., 1964
 Aron M., 1943
 *Arthur J. A., 1992
 Asar T., 1902
 Asch R. 1866
 Askew H O 1912
 Asser 2009
 *Asser R., 2011
 *Associated Electrical
 Industries Ltd.,
 1709
 Asundi R. K., 1733
 *Atkinson R. L. 1906
 Atwood C., 1867
 Atwood F. C., 2009
 Aubert J., 1726
 Ando-Canotill C. B.,
 1903
 *Auer L., 1839 1840
 Aufhäuser D. 1972
 Aunis A. 1907
 Auklpa, J. 1735
 *Austin M. M., 1994
 *Autogewerk Satus
 G m b H 1713
 *Automatic Appiance
 Co The 1709
 v Auwers K. 1818
 v Auwers O 1751
 *Avery, W. & T. Ltd.,
 1713
 Ayar, C. S. R., 1965
 Baar H 1894
 *Babcock & Wilcox,
 Ltd., 1976
 Baborovskij J. 1714
 Bach H. 1926
 Bach K. 1867
 Barharach A. L. 1942
 Bachrach F. 1870
 *Barker C. B. 1744
 Backlin E. 1847
 *Baddley, J. 2003
 *Bader W. 1843
 Badertscher A. F.
 1940
 *Badger C. 1962
 Badger R. M. 1737
 Bailey P. M. 1714
 Bailey G. H. 1728
 Bailey J. R., 1830
 Bailey, K. R. 1867
 Rajas P. 1903
 *Bakelite Corp. 1844
 2012, 2013
 Baker, B. M., 1912
 *Baker, T. T., 1743,
 1749
 Bakhturov V. G., 1953
 Baldwin I. L., 1867
 van Balen, C. L., 1969
 Balla, A. K., 1844 1849
 *Bancel P. A. 1712.
 Bancroft, D. H., 1942
 Bancroft, W. D., 1864
 Banerji S. N., 1722
 Bangs, H., 1831
 Barabachoff, N., 1746
 Barckmann C. 1970
 Bargellini, G., 1809
 Barnitz, M., 1884,
 1892
 Baril O. L., 1813.
 *Barmer Maschinen-
 fabrik A. G., 1903
 Barnes V. E. 1765
 *Baron C. V. A. P.,
 1999
 Barr G., 1724
 *Barzait S., 1963
 Barrenschien H. K.,
 1889
 Barrett, C. S., 1781
 Barro, G., 1827
 Barthel C., 1934
 Barthmeyer, H., 1858
 Bartholomew R. P.,
 1871
 Baskin, N., 2014
 *Bassett H. P. 1996
 Bassett J. R. 1902.
 Bates J. R. 1737
 Batro L. F. 1713
 Batlige 1507
 Bats G. 1932
 Bauer F. C. 1912
 Bauer, K. H., 1800
 Bauer L. H., 1765
 Bance, O., 1751
 Bauer, W. C. 1942
 Bauman, R. W., 1965
 Baumor H. W., 1728
 *Beach A. W. 1923
 Becher, E., 1901
 Bechtel E., 1867
 *Beck C., 1710
 *Beck C. J. 1710
 Beck W. 1830.
 Becker H. 1886
 Becker R. 1923
 Berker R. B. 1713
 Beckman H. 1896
 *Beckmann H. 2021
 Beesley E. 1736
 Bell H. D. 1928
 *Belton Iron Works
 1996
 *Bemberg J. P. A. G.
 1994 2004
 Benavet F. 1782
 Benedict H. 1894
 Benford F. 1736
 Bengtsson N., 1934
 *Benner, H. P., 1943
 *Bennett, N., 1841
 *Bennett, S. C., 1963.
 Benoit, G., 1904.
 *Bensa, P., 1843
 *Benmann H.
 (Firma), 1946
 Benton, T. H., 1932
 *Benton W. A., 1711
 Bercovitch, M. R.,
 1756
 *Berdell, T. V. D.,
 1953
 Berg W., 1894
 *Bergel S., 1950
 Berger, E., 1919
 *Bergl, K., 1925.
 Berl E., 1924 1989
 Berkner, J. F. T., 1931
 *Berthelau, J., 2009.
 *Berlyn J. A., 1792
 Bernan H. 1765
 1767, 1769
 Bernad, A., 1943
 Berner F., 1806
 Bernowis M. W. van
 1967
 Bernhauer K., 1807,
 1812.
 Bernstrom H. O. 1756
 *Bernigan J. B., 1712
 *Bernigan J. J., 1712
 Berry P. A., 1946
 Bert L. 1817
 Bertelot A., 1865.
 Berthelot C., 1967
 1972
 Bertho, A., 1805 1863.
 Berthold G., 1870
 Bertram, F., 1906
 Bertrand G. 1889
 *Bennard A. A., 2017
 Besonova A. 1867
 Better, E. I., 2013
 *Betterson J. O., 1791
 *Beuthner K., 1976.
 Bry L., 1761
 Beyer A., 1727
 Brythorn A., 1921
 Bhagavanam, S., 1735.
 Bhatia L. S., 1721
 Bhattacharya R., 1801
 *Biehl K. 1966
 Bierhalter W., 1965.
 Berry H., 1847
 Bigelow W. D., 1714
 Bigwood E. J., 1723
 Bijvoet J. M., 1716
 Birkens B. S. 1901
 Bilger L. N., 1872
 Biltz, H., 1802
 Biltz M., 1747
 Bismachas 1937
 Bingham R., 1730
 Bingold K., 1887
 Binks, K. R., 1786
 Burkhaug K. 1997
 Burkraut H., 1903
 *Burlay Iron Co. Ltd.,
 1709

- *Burwell A. Wm., 1997
 Bishop, L. R., 1911
 Bishop, P. C., 1911
 *Bjerregaard, A. P., 1946
 *Blache H., 1712
 Blake, E. S., 1523
 Blamberg, I., 1739
 Blanchetière, A., 1433
 Blank, F. W., 1707
 Blanke, K., 1797
 Blatt, A. H., 1820
 Blenkinsop, A., 1739
 Bleyer, B., 1913
 *Bleyer, C., 1976
 Block, F. F., 1423, 1823
 Blumenfeldt, F., 1907
 *Blundell Bros. (Luton) Ltd., 2007
 *Blundell, H. A., 2007
 Blythe G. E. K., 1949
 *Blythe, H. B., 1791
 Boas, I., 1877
 Bodnar, J., 1879, 1889
 Bodroux, P., 1838
 de Boe Z., 1836
 *Böhm, O., 1742
 *Böhme, H., Th. A. G., 1842
 *Böhmer, W., 1791
 *Boeniger, M., 2003
 *Bönemann, P., 1794
 Bornstein, R., 1778
 Böttcher, F., 1723
 Bohn, H., 1483
 Boller, R., 1907
 Bolliger, A., 1836
 Bollinger, G. M., 1723
 *Bollmann, H., 2019
 Bomskov, C., 1885
 Bonath, R., 2017
 *Bond, J. E., 1996
 Bonus, A., 1944
 Bonnet, R., 1881
 Borchardt, W., 1845
 Borgström, L. H., 1715
 Borgstrom, P., 1980
 Bornfeld, F., 1779
 Bornstein, A., 1888
 Bosse, H., 1712
 Bottazzi, P., 1879
 Bouchara, E., 1916
 Bouehard J., 1738
 Bouchetal de la Roche, 1756
 Bouman, J., 1716
 Bourcet, P., 1909
 *Bourdonnet, M., 1943
 Bourdoul, C. (Mlle), 1835
 *Bourgeois, P. L. E. V., 1995
 Boutaric, A., 1738
 Bowen, N. L., 1713
 Boyd, E. M., 1869
 Boyd, J. H., Jr., 1761
 Boyle, J. F., 1949
 Boyle, R. M., 1739
 Bracht, R., 1909
 Bradford R. H., 1776
 Bradley, W., 1775
 Braekken, H., 1718
 Bragg, W., 1957
 Bragg, W. L., 1751, 1766
 Brammell, A., 1772
 Brana, W. A., 1911
 *Bramwell, I. L., 1709
 *Brancart A., 1962
 *Brancart Y., 1962
 Brandt, 1970
 Brandt - Bransmont, Y. (Mme), 1890
 *Brandt, H., 2016
 Brault A., 1831
 v. Brauns J., 1511, 1812
 *Brazay, Z., 1959
 Bredig, G., 1753
 Brenner W., 1934
 Brice, A. T., 1858
 Bridel, M., 1833, 1872
 Bridgeman O. C., 1940
 *Briner P., 1953, 1956
 *Brigando, J., 1937
 Briggs, H., 1967
 *Brightman, R., 2002
 Bril, P., 1904
 *British Celanese, Ltd., 1943, 1963
 *British Dyestuffs Corp., Ltd., 2003
 *British Talking Pictures, Ltd., 1938
 *British Thomson-Houston Co., Ltd., 1709, 1743
 Broadhead, C. F., 1963
 Brockman, C., 1916
 Broglie N., 1739
 Bromley, S. W., 1941
 Brown, G. N., 1852
 Brown, C. P., 1866
 *Brown, E. B., 1922
 *Brown, G. A., 1948
 Brown, H. R., 1914
 Brown, I. A., 1772
 *Brown, J. I., 1963
 Brown, P. E., 1932
 Browns, C. A., 1714, 1931
 Browne, F. L., 2009
 Brownlie, D., 1969
 Bruce, J. A., 1902
 Bruche, R., 1927
 Bruggeman, D. A. G., 1728
 Bruhn, M., 1859, 1889
 Brull, L., 1853
 Brunel, A., 1837
 Brunner, M., 1726
 Brunner, W., 1796
 Brunn, A., 1931
 *Brunn, H., 1931
 Brunschwig, 1979
 Bruynghw, R., 1893
 Buch, A. E., 1960
 Buchholz, J., 1853
 Buchman, J., 1898
 *Buchner, M., 1935
 Buckley, H. E., 1719
 Budnikov, P. P., 1771
 Buchner, T. F., 1723
 Böh II., 1859
 de Buco, V., 1937
 Burger, M. J., 1704
 Buess W., 1778
 Bujard E., 1854
 Bull II., 1721
 Bulas D. H., 1919
 *Bumke H. A. Ges., 1712
 *Bunce F. H., 1790
 Bunn F. S., 1779, 1786
 Burek, P., 1816
 Burgers J. P., 1901
 Burnet F. M., 1463
 Burnell, F. H., 1733
 Bushill J. H., 1917
 *Butcher C. H., 1710
 Butterfield C. T., 1929, 1930
 *Butta, D. C., 2012
 Buylla, B. A., 1946
 Cabanes P., 1919
 Cabot, H., 1999
 Cadden, J. F., 1900
 *Caffrey W. H., 2007
 *Cambridge Instrument Co., Ltd., 1711
 *Camerota, L. A., 1790
 Campanacci, D., 1904
 Campbell, M. A., 1725
 Campbell, W. G., 1714
 Campbell, W. P., 1952
 Canale, E., 1863, 1919
 Canaye, J., 1919
 Cano J. Y., 1722
 Cannavo L., 1899
 Cantilo, E., 1899
 Cape, R., 1912
 *Carborundum Co., 1963
 Carey, T. N., 1892
 Carlo, R., 1906
 *Carlsberg, P., 1741
 *Carlsedt, R., 1713
 Carpenter, D. C., 1896
 *Carpule, G. m. b. H., 1711
 Carrell, A., 1852
 *Carreras R. S., 1769
 *Carrington P. G., 1790
 Carter, B. B., 1858
 Carter, R. H., 1942
 de Carvalho, A., 1912
 Casaburi, V., 1912
 *Casale-Sacchi, M., 1711
 *Cases, M., 1926
 Casper, W., 1891
 Cassan, H., 1941, 1962
 *Cassella, L. & Co., Ges., 2002
 Castiglioni, A., 2009
 Castille, A., 1877, 1882
 *Caven, T. M., 1965
 Caz, A. K. B., 1931
 Chamnade, R., 1852
 Champetier, G., 1752
 Chandler, C., 1853
 Chapman, P. J., 1912
 Chapman, T. S., 1722
 *Chapman, W. D., 1976
 Chappell P. L., 1783
 Chappell, P. L., 1917
 Charaux, C., 1872
 Chargaif, P., 1967
 Charonnet R., 2013
 Chasoff, J., 1899
 Chaudron, G., 1753
 Chauvenet, L., 1753
 *Chem. Fab. Grünau Lamishoff & Meyer A. G., 1709
 *Chem. Fab. von Heyden A. G., 1841
 *Chem. Fab. Oker & Braunschweig A. G., 1943
 *Chemische Fabrik vorm. Sandos, 1841
 Chen M. Y., 1902
 Chernoshkov, N. I., 1931
 *Chester T., 1711
 *Chevalet, P., 1923
 Chevalier A., 1931
 Chevaller, J., 1918
 Chevenard P., 1782
 Chiabrera, S., 1858
 Chateillon, A., 1848
 Chubb, A. C., 1873
 Chichabala, A. E., 1832
 Childs, L., 1941
 Chinn S., 2000
 Chinn A., 1950
 Chirkin, B., 1979
 Chishkov, D. M., 1776
 Chopra, R. N., 1906
 Choudhury, S. R., 1721
 *Christiansen, B., 1790
 Christie, W. A. K., 1933
 Christopher, C. P., 1777
 *Church J. W., 1937
 Crocatté, V., 1861
 Clark, A. J., 1775, 1883
 *Clark, K. A., 1946
 Clark, P. C., 1759
 Clark, R. P. D., 1752
 *Claude Neon Lights, Inc., 1745
 Claus, F., 1932
 Claus, W., 1785
 Clausen, P., 1716
 Clavel (Mme), 1891
 Clegg, E. L. G., 1944
 *Clemens, W., 1973, 1974
 *Clifford, J. J., 2011
 Chmenko, D. R., 1884
 Clouse, J. H., 1718
 Cluzet, J., 1737
 Cobb, J. W., 1969

- *Cochran R. S., 1731
 *Cochrane Corp., 1712
 Cocking, T. T., 1947
 *Cockram C., 1776
 Codomo, A., 1854, 1892
 Coe, M. R., 1915
 Cochill, R. D., 1857
 Cohen, P. L., 1819
 Cohn, B. E., 1728, 1754
 *Cole H. C., 1906
 Cole H. L., 1910
 Coleman G. E., 1965
 *Colias & Altman Corp., 2008
 *Colias, V. A., 1711
 *Colton, J. H., 1963
 *Colour Soapworks (1929), Ltd., 1741
 Combes, T., 1855
 Connel, M., 1873
 *Commercial Alcohol Co., Ltd., 1992
 *Compagnie française pour l'exploitation des procédés Thomson-Houston, 2012
 *Compagnie française de transformation métallurgique 19.4
 *Compagnie Inter nationale pour la fabrication des essences et pétroles, 1923
 *Compagnie des mines, forges et aciéries de Vitoria, 2011
 *Compagnie provinciale de Raffinage de sucre gras sucrés, 1953
 Compromis D., 1855
 Conard, C. K., Jr., 1740
 *Cones, C., 2017
 *Congco, L., 1811
 Conner, R., 1906
 *Consortium für Elektrochemische Industrie G. m. b. H., 1744, 1843
 Conna, C., 1878
 Copeman, P. R., v. d. R., 1874
 *Cortale, H. N. M., 19.4
 Cordouan, R., 1777
 Cornejo A., 1926
 *Corra Products Refining Co., 2018
 Corrie G. R., 1707
 *Corson, M. G., 1732
 Coslow, A. L., 1764, 1765
 *Courtis, Ltd., 1994
 Courtney, O. K., 1941
 Coward, H. P., 1977
 *Cruz, C., 1922
 *Cruglak Chemical Co., Ltd., 2011
 Crane, M. M., 1929
 Crawford, A. M., 1900
 Crawford, W. P., 1762
 *Cremas Marketing Co., Ltd., 1922
 Cremer, E., 1751
 Cremer, H. W., 1752
 Cromwell, H. W., 1896
 *Cromwell, J. C., 1791
 Crookshank, H., 1766
 *Crowley A., 1712
 *Crowther, P. K., 1959
 Cummings, G. D., 1846
 Cury, L., 1902
 Curt, A. L., 1722
 *Curry, W. H., 1789
 *Curry, L. P., 1964
 *Curry, M., 1964
 Cuthbert, P., 1913
 *Dacres E., 1794
 Dalt, P. S., 1857
 Dagnest, E. L. K., 1919
 Dahle, C. D., 1917
 Dale, A. J., 1941
 *von Dalters-Wegner, R., 1709
 Daly, B., 1828
 Dantonovskaya, A., 1855
 Daserfeld, S. J. E., 1711
 Damiro S., 1919
 Dammaville, P., 1881
 *Darco Sales Corp., 1945
 Darcy, M. M., 1940
 Darlot S. P., 1806
 Darrow, P. L., 1728
 *Darrow, V. M., 1794
 Das-Gupta, T., 1754
 *Dassonville, M., 1994
 Daupt, J., 1779
 Davery, W. J. C., 1929
 David, 1726
 Davidson, S. C., 1767
 Davidson, W. P., 1778
 Davidson, W. M., 1729, 1940
 Davies, C. W., 1739
 *Davies, R. O., 1712
 *Davies, W. S., 1789
 Davis, R., 1707
 *Davis, R. H., 1923
 Day, E. M., 1856
 Dean, C. A., 1941
 *Dean, J. N., 2020
 De, P., 1966
 DeLerg, 1854
 Delano, G., 1934
 DeMonte, L., 1729
 Davis, R., 1707
 Deem, J. W., 1924
 Dehinger U., 1717
 Deha, W., 1819
 Delaby, R., 1829, 2015
 *Delbac, A., 2008
 Delvance, (MSe.), 1723
 *Dénos, M., 1774
 Denison, W. P., 1934
 Dent, R., 1714
 Derra, E., 1907
 Desbied, L. B., 1714
 von Deschander, J., 1846
 von Deret D., 1851
 Deslokes, V., 1856
 *Deutsch - Europäische Quarzschmelze Ges., 1963
 *Deutsche Gasflicht-Act - Amer. Ges., 1925, 1942, 1964
 *Deutsche Gasolin A.-G., 1925
 *Deutsche Gold und Silber - Schmelzwerk vorm. Rönner, 1791, 1943, 1954
 *Deutsche Hydrierwerke A.-G., 1925, 1927, 1953
 *Deutsche Luftver-Baugesellschaft m. b. H., 1709
 Devet Kijeva, M. D., 1912
 De Voord, 1909
 Deacon, P., 1918
 Dick, J., 1755, 1909
 Dickson, P., 1707
 *Dickson, S. J., 1953
 Dieckmann, 1794
 Dieert, P., 1927, 1930
 Dieterle, R. K., 1810
 Dieterle, W., 1745
 Dieterle, P., 1979
 Diets E. M., 1918
 *Diets, H. P., 1799
 Diestricher - Maiste 1891
 Dingemans E., 1837
 Dinmore, K., 1810
 Dirk, B., 1935
 Dieckendorfer O., 1825
 *Ditz/Ders Co., Ltd., 1943
 *Dittmar, R., 2021
 Dixon J. K., 1913
 Dixon, P. J., 1917
 *Doell, H., 1941
 Doggett, R. A., 1769
 Dojms, S., 1926
 Dolmage, V., 1770
 Donan, J., 1747
 *Donner C. F., G. m. b. H., 2019
 Dornbachka A., 1721
 Dove, W. H., 1907
 *Downman A., 1794
 *Dougherty, W. P., 1964
 Drake, E. T., 1912
 Drake, T. G. H., 1909
 *Draper, W. P., 1994
 *Dreyfus, C., 1993
 1964
 *Dreyfus, H., 1829, 1943, 1953, 1993, 1995, 2009, 2007
 Dreyssing, C., 1925
 Driggers, R. P., 1930
 Duce, J. G. P., 1871
 Dymars, P., 1723
 Drumm, P. J., 1879
 Drummond, J. C., 1876, 1882
 Dubais, J., 1831
 Dubois, G. H., 2020
 DuBois, L., 1949
 Dulose, 1936
 Duloway, R., 1724
 Dubal, L., 1771
 *Duramp, A. J., 1956
 *Duckham, A. M., 1925
 Duckworth, J. P., 1949
 Ducas, J., 1728
 Duda, P., 1872
 Dürr, P., 1728
 *Dufay, L., 1745
 Dufort, A., 1899
 Dufosse, 1972
 Dugand, J. B., 1802
 Dular, C., 1766
 Duncan, D. R., 1752
 *Dunham C. A., 1713
 *Dunlop Rubber Co., Ltd., 2020, 2021, 2022
 Duro, B. W., 1997
 *Duns, J. S., 1955, 1956
 *Dusch, O., 1709
 *Du Pont de Nemours, E. I., & Co., 1844, 1953, 1954, 2023
 Dupont, R., 1740
 *Durr, P., 1729
 Duval, C., 1749
 Duval, C. (Mme.), 1749
 Dvortakha - Barsheva, K. M., 1907
 Dvorak, E., 1743
 Dye, W. B., 1707
 *Eagle Pencil Co., 1953
 Easterfield, T. H., 1892
 Eastman, V. J., 1900
 Easton, R. P., 1707
 Eastwood, A. H., 1909
 Eble K., 1916
 Eckart, C., 1729
 von Eckermann, H., 1767
 Eddy, N. B., 1903
 Edin, H., 1921
 Edin, B., 1734
 Edin, J. T., 1845, 1846
 *Edwards, R. & Co., Ltd., 1743
 *Edwards, W. R., 1743
 Efron, J., 1905
 *Egenfeldt-Olsen, P. E., 1709
 Engelhart, E., 1907
 Eloff, G., 1796, 1778

- [illegible]

- Gonyer F. A., 1767
1768, 1770
Good, R. C., 1781
*Goodell C., 1966
Goosmann J. C., 1932
Gorensou E. A., 1767
*Gordon, C. W., 1709
Gordon M., 1866
*Gorlich F., 1963
Gorman A. E., 1930
Gottner, R. A., 1721
Gosman B., 1978
*Goss B. C., 2000
Gossner B., 1763,
1770
Goudrey R. F., 1928
*Goudelock W. B.
Q. B. 1793
Goudriaan F., 1729
Goversers J., 1891
Goyle, D. N., 1738
Gräf, G., 1863
de Gracve, P., 1857
Grate, V., 1924
*Grah F., 1791
Graham R. P. D.,
1767
Grandjean E., 1901
*Graschoff E., 1712
Grassmann W., 1870
*Grasol, O., 1922
*Gray, C. H., 2022
Gray, I., 1832, 1902
Gray, R. M., 1852
Grebe, L., 1746
*Green, A., 1996
*Green, A. J., 1713
*Green, B. E., 1713
*Green, W. H., 1931
*Greenhalgh, E., 2008
Greenstein M. L.,
1909
Gregory, J. W., 1770
Griebel C., 1946
Griffin, C. W., 1708
*Griffin J. H., 1963
Griffing E. P., 1762
Grimmer, W., 1916
Grinfeld R., 1733
Grip, E., 1766
*Grohmann H., 1958
von Gronow, H. E.,
1721
Gross, J., 1707
Grosscup C. G., 1751
*Grote E. R., 1710
Grote G., 1745
Grove Palmer, F.
1788
Guastardo G., 1901
Gubelmann, I., 1821
Guedon A., 1808
Günther C., 1959
Günterschulze A.,
1730
Guertler W., 1784
1935
Guilford, R. S., 1959
Gumaraes D., 1765
Gulbas, F., 1730
Gumz W., 1968
Günther P. R., 1876
Gurwitsch A., 1851
*Gutckoßnung
shütte Oberhausen
A. G., 1744, 1844
*Guttenberger, M.,
1968
Gutzalt, A. M., 1891
H. J., 1902
*Haake R., 2018
*Haake W., 2018
Haas M., 1784
Haberlandt L., 1892
Hagg G., 1718
*Härden, J., 1793
*Härdler F., 1790
*Hagemann H., 1792
*Haglund E., 1992
Haglund E., 1914
Haba P. L., 1725
Haid A., 1994
*Hain V. A., 1712
Halberstadt, O. T.,
2015
Hale M., 1738
Haley, D. E., 1942
*Hall H. C., 1791
Hall, J. C., 1967
Hall J. L., 1910
Halla P., 1764
Haller P., 1944
Hamel C. R., 1774
Hammarsten G., 1961
*Hammond C. F.
2008
Hammond, H. R.,
1707, 1922
Hank A., 1918
*Hancock L. J., 1701
*Handwerk, E. C.,
1790
*Hanna C. R., 1958
Hanna, G. D., 1838
Hanna W. F., 1873
Hanneke, P., 1746
*Hansatische Ap-
paratbau Ges.
vorm. L. von Bre-
men & Co., 1925
*Hansen, C. J., 1975
Hanson, M. H., 1808
Harrang L., 1718
Hardy, V. R., 1709
Harkins W. D., 1778
Harmen S. W., 1941
*Harpun J. H., 1713
Harrington V. F.,
1708
*Harris, H., 1790
*Harrison C. F. R.,
1974
Hart L., 1949
Hartek F., 1718
*Hartel H., 1726
*Harter H., 1711
*Hartley E., 1931
Hartman F. A., 184
Hartman F. V., 1784
Hartmann, B. G., 1914
Hartmann, E. C., 1784
Hartmann, W., 1914
Hartweg, H., 1918
Hartwell F. Z., 1941
1942
*Haselmann H., 1993
Hasenfratz V., 1949
Haslam, R. T., 1942
Hasselberg H., 1874
Hasselström, T., 1823
*Hassmann, F., 1906
Hatzfeld, W. D., 1929
Helmaker P., 1771
Hauber, E. S., 1815
*Haudorfer E., 1845
Hauner, E. A., 1722
Hannemann, W., 1770
*Hawerlander A.,
1966
Hayashi T., 1824
2000
Hayashi Y., 1937
Hayes, A. O., 1772
*Hayot, R., 1711
*Hazard Flomond M.,
1711
Hebestreit O., 1961
Hedgkatti R. M.,
1914, 1938
Heres P. A., 1894
Hegel, K. T., 2001
*Hedderberger Gela-
tine Fabrik Steuss &
Co., 2005
Heil J. W., 1728
Heibergster, R., 2014
Heipens S., P., 1821
Heinrich F., 1935
Hein, H., 1763
Hensenberg, E., 1747
Held I. W., 1899
Helfmann, A., 1795
Helmann L. M., 1911
Helmström, H., 1876
Henderson D., 1968
Henderson, G. H.,
1731
Heurtenberg J., 1718
Heurtenberg O.,
1779, 1782
*Henke C. O., 2012
Henkel J. S., 1921
*Heinkel & Cie G. m.
b. H., 1942
*Heule F., 1843
*Henley's W. T. Tele-
graph Works Co.,
Ltd., 1745
Henne A. L., 2020
Hennig N., 1907
*Heutzel F., 1742
1743
*Heurats W. C., G. m.
b. H., 1792
Herd C. W., 1914
Hermesdorf H., 1746
*Hermesdorf Schom-
burg Isolatoren
Ges., 1743
Hermann R. E. S.,
1765
*Herold A. G., 2012
Heron A. M., 1964
2009
Herr E. F., 1822
Herr R., 1772
Hernick T. L., 1929
*Herrig G. E., 1938
Herschman H. K.,
1778
Hertel, E., 1720, 1825
Herty, C. H., Jr., 1777
*Hersfeld E., 1791
*von Hertke, G.,
1963
Hess R., 1836
Heutekian H., 1928,
1929
Heulin, F. F., 1872
v. Hevesy, C., 1751
*Heylandt, C. W. P.,
1711
*Heyna H., 2004
*Heynemann L., 1966
Heyrovsky, J., 1978
Heywood H., 1714
Hibbert H., 2010
Hicks, C. G., 1920
Hicks, C. S., 1822
Hiddeit T. P., 1800
1801, 1882, 2014
*Hilger, C., 1968
*Hill, J., 2003
Hill, J. H., 1948
*Hill R., 2012
Hilgruber K., 1853
Hilg, F., 1914
Hilward O., 1948
Hind, H. L., 1844
Hind, S. R., 1960
Hinglais, H., 1801
Hinselman U., 1858
Hirano S., 1819
Hiron, J. (Mlle.), 1929
Hirsch, A., 1834
Hirsch, M., 1717
Hirschhorn, L., 1902
Hirsch Kaufmann, H.,
1906
Hirschkind, W., 1752
Hirschmann, H., 1820
*Hirt, H., 1713
Hiltzberger K., 1853
Hittig W., 1950
Hjort A. M., 1855
Hloch A., 1924
Hoag L. E., 1757
Hoare B. G. W., 1949
Hobson G. V., 1959
Hochrein M., 1858
Hodges J. H., 1727
Hodgkiss H. E., 1942
Höber R., 1906
Högbom A., 1764
Högler F., 1906
Hölder F., 1805
Hoem, E., 1897
*Hönd H., 2012
Hönigsmid, O., 1715
Hoffmann P. G., 1972
Hofman-Bang G.
1944
*Hofmann F., 1955
Holden R. J., 1964
*Holland W. W., 1956
Holman B. W., 1774
Holmes A., 1729
*Holmes C. W. H.,
1709
*Holophane, Ltd.,
1709

- *Holroyd R. 1914
Holt, E. C. 1917
*Holt, T. W. 1942, 1943
*Holtz A. 1955
Holzer, H. 1757
*Hollindustrie-Werke Josef Benz, A. G. 1966
Hollverkohlungsges. A.-G. 1756
Hommon, C. C. 1929
Honda K 1716
Hood Rubber Co 2021
*Hooley W. C. 1700
*Hooley L. J. 1945
*Höppe Heating & Lighting, Ltd 1713
Hopper T. H. 1932
*Horn S. 1954
Horn G. 1947
Horrall, H. P. 1918
*Horst, W. P. 2022
Hoskins, R. G. 1847
Houben, J. 1822
Houser, G. C. 1925
Houls, R. C. 1795
Howard, P. P. 1935
*Howard, W. R. 1941
*Howden, J. & Co., 1712.
*Howe, A. J. 2021
Howell, S. F. 1855
*Howes, D. A. 1934
Howland, L. H. 1915
Hoyt, A. 1909
Hoyt, L. F. 1759
Hruba F. 1921
Huang, K. K. 1903
*Hubert, C. A. 1711
Hucker, G. J. 1596
Hockett, H. C. 1911
Höfner, F. 1805
Huggins, M. L. 1805
Hughes, T. A. 1493
Hugill, W. 1960
*Hume, J. H. 1712
Hummutzsch, W. 1750
Humphrie, F. H. 1739
Hunt, P. S. 2018
*Hunt, W. P. 1713
Hunter, G. 1852
Huntzinger, M. E. 1888
Huppert, O. 1970
Huppmann, 1761
Hurd, C. D. 1814
*Hutley, T. F. 1976
*Huron Industries, Inc. 1713
Hussey, R. E. 1900
Huston, H. A. 1714
*Hutchinson, H. B. 1945
Huter, J. 1753
*Hytrochuk, G. m. b. H. 2021
Hylleraas, E. A. 1734
*t G Farbemundstrie A.-G. 1744, 1745, 1749, 1789 1791
1794 1834 1839
1810 1841, 1812,
1811, 1845, 1928,
1913 1920, 1934
1953, 1954, 1957,
1954, 1964, 1973,
1974 1975 1944
1945, 1956, 1992
1993, 1991 1995
1996, 2001, 2002
2003 2004 2005
2006, 2007 2009
2011 2012, 2016
2017, 2018 2019,
2020, 2021, 2022
Ievit A., 1776
Himku V. P., 1952
Ilkoff I. 1903
Ilies 1970
Immke tl 1941
*Imperial Chemical Industries Ltd., 1711,
1712 1839, 1841,
1843 1811 1845,
1953, 1955, 1956,
1974, 1944, 1993,
1996, 2002, 2006
2011, 2012, 2013,
2020
*Impreva Holzimprägnierung und Holzverwertung A.-G. 1966
*Inhausen, A. 2017,
Inoue, Y. 1800
Inouye, J. M. 1857
*International Commission Engineering Corp. 1712
*International Nickel Co, The, 1791, 1793
*International Patents Development Co., 2016
Ionnidis Z. M. 1949
Iselovskii, I. V. 1953.
*Ipson, C. L. 1744
Iredale, T. 1736
Irye, T. 1952
Ishwara, T. 1727
*Israel, F. 2013
*Italica (Società anonima), 1712
*Ito, C., 2016
Ivey, J. W. 2000
Ivy, A. C., 1901
Iwamoto, Y. 2019
Izume, S. 1678
Jachael, J. 1965
Jackson, A. 1780
Jackson, F. H. 1965
Jackson, H. A., 1927
Jackson, V. T. 1707
Jackson, W. W., 1766
Jacobs W. A., 1835
Jacobsohn, K. P. 1850
Jacobson, C. A. 1810
*Jacobson, D. L. 1974
Jacoby, M. 1900
Jadassohn, W. 1903
Jadin, J. 1864
Jarvinen K. K. 1714
Jandeleur W. 1944
1959
Janot, M. M., 1822
Janson A. A. 1899
Jansson, G. 1871
Jaretsky R. 1814
Javiller M. 1847
Jeghers H. J. 1857
*Jenkins W. J. 2011
Jennings T. F. 1781
Jensen, B. N. 1756
Jensen, H. L. 1931
Jochims J. 1845
Joffe, N. 1907
Johansson D. 1992
Johnson, C. M. 1756
Johnson P. 1781
Johnson F. D. 1999
Johnson J. B. 1780
Johnson J. M. 1904
Johnson J. R. 1795, 1825
*Johnson J. V. 2021
*Johnson L. S. 1996
Johnson T. B. 1830
Johnson W. C., 1724, 1731
Jones, B. M. 2000
Jones, P. W. 1928.
Jones, G. 1725
Jones, G. W. 1997
Jones, H. A., 1939, 1910
Jones, H. C., 1774
Jones, H. F., 1929
Jones, W. C., 1935
de Jongh, S. F., 1845
Jono, Y. 1846
Jonsson, B. 1723
*Jordan, P. 1792
Jordan, H. J., 1890
Jordt, H. 1952
*Joseph P. L., 1709
Joslyn, R. G. 1926
*Jos-Pe Farlenphot G. m. b. H. 1149
Jouan, S. 1909
Jourdan, A., 1961, 1962
Joy, W. F. 1725
Junt, R. E. 1966
Janzeblut, C. W. 1909
*K. D. P., Ltd., 2020
Kabanov, D. N., 1740
Kahn, R. L. 1896
*Kaiser, R. 1791
*Kaiser Wilhelm Institut für Eisenforschung 1709
Kaku, T. 1945
Kalantar, N. 1950
Kalb, G. 1762
*Kaherky, J., 1978
Kafankova, M. N., 1833
Kallor, P. 1693
Kalmán, P., 1806
Kamiya, T. 1882
*Kamm, E. D. 1974
Kamosta, Y. 1934
Kane, H. P., 1890
Kanco, H., 2000
Kaputras, H. J. 1757.
Kar, H. A. 1729
Karaoglanov Z. 1750
Karavayev N. M., 1941
Karczag, I., 1909
1910
Kareh, A. 1879
*Karolus A. 1710
*Kärpälä J. 1945
Karrer P. 1990
Karrer P. 1791 1876
*Kary, O. 1915
Katzhavin N. A., 1831
Kassensky B. 1947
Kashihara H. 1726
*Kastner R. 1996
Kato T. 1951
Kate J. R. 1732
Kawakami M. 1727
Kay A. A. 1866
Kayser P. 1847
Keefler C. S. 1903
Keeper C. P. 1927
Kehi R. 1912
Kehren 2011
*Keller J. J., 1713
Kellaway, C. H., 1904
*Keller C. L. 1996
Keller, P. 1937
Keller, J. P. 1755.
Kent-Jones, D. W., 1914
Kerridge, P. M. T., 1862,
Kerran, W., 1782,
Kertész Z. I., 1919
Khan, S., 1861
Khokhlov, V., 1950
Kidani, Y., 1778,
Kiefer, P. J. 1925
Kieffele, P., 1918
Kiliani, H., 1834
*Kilham, G. B., 1713
*Kilten, D. J., 1937.
Kimura, H., 1860,
Kändler, K., 1810,
King, F. R., 1923,
King, H. H. 1910
King, H. S., 1796, 1799
King, R. H., 1709.
*Kinsley, C., 1711.
*Kinsel, A. B., 1792
*Kirkup, R. H., 1711.
Kirschbaum, E., 1921
*Kirschbraun, L., 1931,
*Kirschbom, G. N., 1701.
Kies, S. A., 1979
Kitaato, Z., 1624
*Kittel, T., 1792,
Klabunde, H. K., 1826
*Klabr, P., 1712,
*Klavich, W., 1844
*Klein, A. B., 1748
Klein, B., 1866
Klein, M., 1885
Kleiner, I. S., 1903
Klemm, W., 1751

- Kline D S. 1899
 Kluke W. 1951
 Klumpp T G. 1898
 Kneust W. 1773
 Kneys, C. 1866
 Knigge G. 2016
 Knipovich Y. N. 1766
 *Knoll A-G Chemische Fabriken, 1844
 *Kowles A E. 1743
 Kox J D. 1777
 Knyale W L. 1789
 Kautern, M. H. 1917
 Kuber, S. 1837
 Korb, F. C. 1861
 Korb L. 1762
 Koch S. 1762
 Kockás J. 1920
 Kodama S. 1852
 Köcher Z. 1937
 Köhler F. 1843 1849
 Kocich C. P. 1823
 Koenen H. 1928
 König W. 1823
 Kocppe, H. 1846
 *Köber, H. 1992
 *Koerver, F. W. 2021
 Kottchau R. 1833
 Kofman, T. 1738
 *Kohlenscheiderung G m h H. 1731
 *Koklmeyer, F. J. 1799
 Kohlshütter, H. W. 1799
 *Kodl, T. 1963
 Kojima, K. 1890, 1900
 Kolb, H. 1787
 Kolke, Fr. 2010
 Kolle, W. 1869
 Kolmer, E. 1754
 Kolsky, S. I. 2000
 Koonstakara, L. 1878
 Kondo, T. 1848
 KooHana, D R. 1873
 Kopfermann H. 1733
 *Koppers, H. A-G. 1775
 Kora, R. 1965
 Korobova, L. N. 1722
 Kosake, S. 1900
 Kosaka, Y. 1971
 Kostrikh, Y. M. 1927
 Kostria, K. 1961
 Kothny, E. 1784
 Kounisky, R. 1904
 1905
 Kuvelovčová, J. 1816
 Kosa F. 1905
 Kracek, F. C. 1718
 *Krankle, G. 1843 2006
 Kranz F. H. 2006
 Kratz, P. 2018
 Kraus E. H. 1769
 Kraus, R. 1869
 Krause, A. 1951
 Krause F. 1934
 Krcul P. 1963
 Kreha, H. A. 1843
 Kreiselman, J. 1890
 Krishnamurti P. 1732
 Krivolek, V. N. 1783, 1785
 Khl A. 1714
 Krowber, L. 1867
 Krowts, C. 1891
 Kroll, W. 1743
 Krow, L. C. 1705
 Kromavakü, A. A. 1855
 Krueger, R. H. 1781
 Krumpel, O. 1770
 *Krupp F. A O. 1712, 1791
 *Krupp P. Gaussonwerk A. G. 1925
 Krayt, H. R. 1720
 Kudryavtsev, N. T. 1740
 Kührt, H. 1854
 Kueka, M. 1811
 Kuhnso, J. 1860
 Kirschner K., 1910
 Kuba, R. 1719
 Kulkarni V. H. 1910
 *Konberger, A. P. 1975
 Kundert, A. 1758
 Kuntz W. 1766
 *Kunze K. 1712
 Kurbatov, I. D. 1781
 Kuroda H. 1903
 Kurotskin, T. J. 1921
 Kuechnarjew, M. A. 1928
 von Kuthy, A. 1831, 1880
 Kusak, J. 1937 1904
 Kusato S. 1845
 Kusnetsow, S. J. 1851
 Labbé, H. 1921
 Laby, T. H. 1730
 Lacian, N. C. 1877, 1895
 Ladenburg R. 1733
 Laeger, H. 1965
 *Laug B. 1885
 Lang J. 1780
 Laksner, A. 1814
 *Lamb A. B. 1844
 La Mer, V. K. 1724
 Lam, R. 1873
 Lammert, O. M. 1725
 Lampert L. H. 1917
 *Lander C. H. 1876
 Landolt H. 1728
 Lange B. 1705 1729 1850
 Lange T. 1967
 Lange W. 1754
 Langacker H. 1948
 Langer J. A. 1893
 Langmuir I. 1731
 Lamewski M. 1880
 Lapsour E. 1885 1887
 Larkum N. W. 1707
 *de La Rochette C. 1973
 Larsen F S. 1769
 Larsen, R. L. 1899
 *Larson, J. M. 1709, 1713
 Lawalle, H. 1892
 *Lastovicks, V. 1713
 Laenderm, J. D. 1764
 Launay, L. 1903
 Laurent, V. (Mlle). 1904
 *Lautenschlager, F. & M. G. m h H. 1711
 Lavin, G. I. 1737
 Lavie, T. P. 1802
 *Lawson W. F. 2012
 Leake, C. D. 1761, 1902
 Leaver E. S. 1775
 Lebedyantssev, A. N. 1935
 Lecon, R. 1773
 *Lederer A. 1739
 Lederer, E. L. 2014
 Lederer, E. R. 1882
 Lee, J. A. 1931
 Lee M. O. 1844
 *Lefsaer, O. 1743
 *Legendre, R. A. 1923
 LeGrays, M. 1967
 Lehmans Facius, H. 1894
 *Lehrke H. 1853
 Leeb, T. 2010
 *Lewer, R. 1712
 Leuter, F. 1762
 *Leflap, O. 1884
 Lemarchanda, M. 1780
 Lemagne, M. 1852
 Lendle L. 1906
 Leonard Jones, J. P. 1719
 Lemox, W. G. 1900
 Lenoble, E. 1728
 Lenhardt A. 1920
 *Leopol Internationale Patentverwertung G m. b. H. 1963
 Leske, E. L. 1898
 *L'Etrenge, W. O. 2012
 Lesler, A. 1831
 *Levan, J. 1962
 Levar, P. A. 1805, 1818
 Levine, B. S. 1895
 *Levy D. L. 2018
 Levy W. Nathan 2017
 Levy S. 1733
 Levy S. A. 2009, 2010
 Lewis W. C. M. 1849
 Liberalis C. H. 1916
 *Libbey Owens Glass Co. 1962
 Lieben F. 1850
 Lieberich E. 1779, 1780
 Lichtenstein L. 1892
 *Lierg F. 1719
 Liezegong R. E. 1747
 Lifschits, I. 1824
 *Lienfeld L. 1993 1994 1995
 Lim C. E. 1891
 Liock, G. 1763
 Linder, E. G. 1717
 *Lindman, E. I. 1866
 Lindquist, F. E. 1738
 Ling S. M. 1875
 Lishorst P. F. 1727
 Linton, P. B. 1714
 Linsal, W. 1846
 Lions, P. 1823
 Liptov S. M. 1722
 Lipets I. M. 1901
 von Lippmann, F. O. 2017
 Lira, J. di L. 1863
 Lison L. 1854
 Litting T. 1893
 Littlejohn, W. 1880
 *Lischer, G. 1943
 Lloyd, B. 1867
 *Lloyd L. N. 1712
 Lychow, M. 1867
 *Lockwood, D. S. 1858
 *Loerelle C. 1962
 *Lodge, W. C. 1997
 Lohb, L. 1902
 Loewy, A. 1888
 *Loftus, F. H. 1791
 Lovinov, N. I. 1982
 Lowmear, J. 1723 1857, 1861, 1895
 London, P. 1715
 *Longworth, J. R. 1713
 Lonsdale J. T. 1784
 Loofstrom J. R. 1878
 *van Loon, C. 1840
 van Loon J. 2013
 Lora y Tamayo M. 1758
 Lorentzen, J. 1978
 Lotze, L. 1750
 Lothrop R. E. 1920
 Lotz, F. 1753
 Louré H. 1716
 Lowenstein, W. 1907
 Lowry, C. D. Jr. 1796
 Loyarte R. G. 1733
 Lobanski, A. 1707
 Lobko, A. 1883
 Lucas H. 1732
 Lucata E. 1977
 Lucker, O. 1796
 Loppo-Cramer 1747
 Lüth P. 1924
 Lomstre, A. 1899
 Lund, R. J. 1772
 Londerkrb, H. 1939
 *Lush, E. J. 1710
 *Lydtin, P. 1966
 Lyons, T. A. 1818
 Maas, E. 1779
 *McAulay, J. 1953
 McCartney, J. E. 1869
 Macchia O. 1760
 McCollum, J. L. 1910

- McCaughy, J. F., 1743.
*McDonald, J. A., 1976.
McEachern, D., 1912.
McElvain, S. M., 1811.
McEwen, G. F., 1714.
*McFarland, J. L., 1741.
Macfarlane, C. M., 1776.
McGregor, A. G., 1775.
Machatschki, F., 1766.
Machek, G., 1829.
*Machon, W., 1793.
Macht, D. I., 1907.
McIntire, J. C., 1940.
McIntosh, D., 1720.
McKeown, M. R., 1774.
McKibbin, R. R., 1939.
McKie, M., 1863.
Mackie, T. J., 1962.
*MacLaren, R., 1715.
*McLean, H. L., 1974.
McLuckie, C., 1995.
MacMahon, P. S., 1721.
*MacMichael, H. R., 1791.
*McNish, A. F., 1962.
MacPherson, R. S., 1781.
Madelung, W., 1826.
*Madge, E. W., 2020.
Madorosky, S. L., 1775.
*Magarawa, K., 1790.
Mayah, M., 1846.
Magath, M. A., 1855.
Maggiore, R., 1915.
Magnusson, M., 1774.
*Mahler, G. T., 1790.
*Mahoux, A. P., 1792.
Matlinder, R., 1782.
Matthe, A., 1795.
*Main, J. A., 1923.
Mann, H., 1705.
Mannner, F., 1858, 1859, 1889.
Mann, J., 1877.
Majumdar, R. C., 1732.
Mahk, K. S., 1893.
Malinowski, V. S., 1831.
Malison, H., 1071.
Malloy, A. M., 1896.
Maloney, E. P., 2009.
*Manchester, C. C., 1709.
Mandrugin, V. N., 1936.
Mangold, E., 1831.
*Manley, R. E., 1966.
*Maancasnarobren Werke, 1790.
Manske, R. H., 1832.
Mantell, C. L., 1740.
Maquelli, C., 1913.
Manzano, H., 1825.
Marelli, R., 1924.
*Marenzi, A. D., 1877, 1878.
*Marques, J., 1934.
Marischka, 1970.
*Marischka, C., 1978.
Mark, H., 1937.
Markaryantz, A., 1950.
Markosov, P., 1940.
Markovich, M. B., 1979.
Marogna, G., 2015.
Marsh, P., 1928.
Marti O. K., 1742.
Martin, G., 2020.
Martin, J. T., 1818.
Martia, J. W., 1932.
Martia, M., 1844.
Martia, O., 1912.
*Martia, W. K., 1792.
Marx, H., 1883.
*Maschinenaufabrik
Buckan R. Wolf A.
G., 1974.
*Maschinenfabrik - An
stalt Hamboldt,
1959.
Masere, M., 1916.
*Masters, F. G., 2022.
Masuda, S., 1900.
Masuda, Y., 1801.
Masuda, C., 1717.
*Matbenas, W., 1792.
Matheu, C. P., 1825.
Matlin, H. A., 1878.
*Matschok, G., 1963.
*Matsumoto, Z., 1709.
Matsumura, S., 1912.
*Matter, O., 1929.
Matthijssen, H. L., 1757.
Matusevich, V., 1973.
Matzner, M. J., 1853, 1902.
Mayer, M. E., 1866.
*Mayer, N., 1935.
Mayhan, A., 1756.
Masur, J., 1716.
*Mead, G. R., 2022.
Medala, L. S., 1867.
Mehl, E., 1766.
Mehl, R. P., 1731.
Mehmel, M., 1768.
Meinel, K., 1933, 1939.
Mewel, K., 1731.
Meiser, W., 1805.
Messner, G., 1665.
*Messner, J., 1929.
*Meissner, T., 2004.
Meissner, W., 1771.
*Melant, A.-G., 1983.
Melchior, P., 1784.
McLean, E., 1926.
Melior, D. P., 1736.
Melfer, W., 1963.
Meugebier, H., 1915.
Menschuthin, B. N., 1714.
Messer, G., 1763, 1765.
Messner, R. C., 1797.
*Metallges. A.-G., 1745, 1790, 1922, 1954, 1975.
*Metals Protection Corporation, 1743.
van der Meulen, J. H., 1757.
Meyer, J., 1965.
Mayer, K. H., 1947.
Meyer L., 1707.
*Mayer Mineral Sepa
ration Co., 1789.
Meyer, O., 1753.
Meyer P., 1979.
Meyerhof, O., 1940.
Meyers, J. L., 1811.
Meyers, V. C., 1857.
Michael, A., 1802, 1803.
Michaux, A., 1876.
Michael, P., 1801.
Michael, H., 1801.
*Michael J. W., 1973.
Michel, A., 1742.
Michel P., 1704.
Michelson, J., 1844.
Migley T. Jr., 2020.
Miehr, W., 1901.
Miford P. L., 1896.
*Miflar, W. J., 1741.
*Mifler C. P., 1966.
Mifler R. J., 1875.
*Miflaken M. G., 1934.
*Mifla, H. A. T., 1953.
*Miflmas, A. G., 1959.
Mifskat Bogdanova, E. T., 1875, 1893.
Mifla, B., 1855.
*Mifscuppe Glass Co., 1962.
Mifschell, A. S., 1714.
Mifschell, H. V., 1933.
Mifschell, J. H., 1938.
*Miftsch, A., 1844.
Miyama, M., 2000.
Miyata, M., 1746.
Mizoshita, T., 1812.
Modderman, J. G., 1720.
Modell, D., 1782.
Modell, D. L., 1764.
Modern, P., 1895, 1896.
*Modi, K., 1935.
Mohl, H., 1960.
Moller, H., 1715.
Mönnig, H., 1965.
Möhlman P. W., 1929.
*Mohr, A., 1790.
Mohr, P., 1720.
Mohr, W., 1916.
Moldenhauer, O., 1835.
Moldenke, R., 1777.
Molnar, E., 1850.
Mouasterio, G., 1853.
*Mouneff, J. Ltd., 1942.
Moochaupt, M., 2013.
*Montan- und Indus
trialwerke vorm
Joh. D. Starck, 1789.
Moore, C. R., 1820.
Moore, E. E., 1826.
Moore, M. D., 1826.
Moore, N. S., 1883.
Moore, R. A., 1911.
Moore, R. R., 1780.
Moreau, 1899.
Morel, R., 1857.
Morey, G. W., 1714.
Morgan, G. F. V., 1919.
Morgan, G. T., 1753, 1809.
Morgan, J. L. R., 1725.
Morrhau, 9, 1857.
*Morrhau V., 1711.
Morkett K., 1929.
Morkett R. S. P., 1857.
Morrill C. A., 1922.
*Morrill J. C., 1944.
Morrill J. P., 1912.
*Morrison J. C., 1957.
*Morrow C. H., 1715.
Morton, C., 1947.
*Morton W. A., 1962.
Mowbray, C., 1894.
Mottorn, H. H., 1920.
Moulton C. H., 1825.
Mourmand, G., 1841.
Mours, J., 1844.
Mourot, G., 1841.
*Moser, A., 1743.
*Muschka, J., 1931, 1926, 1997.
*Mueller, C., 1975.
*Mueller C., 1811.
*Mueller C. H. F. A., C., 1709.
*Mueller, P., 1726, 1751.
*Müller, E., 1963.
Müller, K., 1739.
*Müller, P. H., 1712.
Müller, R. H., 1799.
Müneb, A. P. W., 1847.
*Murdan, M., 1813.
Mukherjee, J., 1721.
Muhno, M. O., 1902.
*Muhaupt P., 1791.
Mund, W., 1877.
Murskaml, J., 1990.
von Muralt, A. L., 1846.
Murnour, H., 1997.
Murgulescu, I. G., 1753.
Mussnug, P., 1770.
Myra, C. N., 1909.
*N. V. de Bataafche Petroleum Maatschappij, 1839, 1842, 1843, 1844, 1975, 1974, 2011.
*N. V. Maatschappij tot Beheer en Exploitation van Oliebronnen, 1962.
*N. V. Machinebouw en Apparaten Fabrieken, 1973.
*N. V. Mijnbouw en Cultuurmaatschappij Doetin, 1986.
*N. V. Nederlandse Technische Handel Maatschappij "Giro", 1715.
*N. V. Philips' Gloeilampenfabrieken, 1710, 1745, 1795.
*N. V. "Solopoi" In rensfabrieken, 1973.

- Exploitation van het
System Polyvus,
1963
- Nagakawa S., 1860
- Nakahara W., 1476
- Nakamura S., 1952
- *Nance C. W., 2016
- Nara M., 1948
- Narayana N., 1871
- Nash H. F., 1907
- Nasmith G. G., 1927
- Natanson L., 1733
- Naumann F., 1729
- Nayar M. R., 1721
- Neale A. V., 1899
- Nebel R. R., 1843
- Needham L. W., 1965
- Negreev V., 1781
- Nehl F., 1782
- Nellensteijn P. J.,
1971
- Neller, J. R., 1919
- Nelson F. K., 1874,
1929
- *Nelson J., Ltd., 1843
- Nelson, R. E., 1617
- Nemes, A., 1935
- Nepp, J., 1802
- Nellmann, F., 2007
- *Neufeld, M. W., 1792
- Neumann K. E., 1990
- *Neville, R. C., 1926
- Neyros, K., 1948
- Newman R. W., 1917,
1918
- Newman, S. H., 1999
- *Newton J., 1962
- *Newton N. B., 1938
- Nayron, C., 1873
- Nicholas W. W., 1962
- *Nichols Products
Corp., 1713
- Nickerson J. L., 1731
- Nicomoff, D. F., 1776
- *Nielsen H., 1965
- Nielsen, N. A., 1909
- *Nielsen, N. J., 1711
- Nifontova, S., 1978
- Nightingale, S. J., 1785
- Niklas, H., 1934
- Nilsson, H., 1870
- Nilsson R., 1847
- Nishida, K., 1870,
1957
- Nishihara, H., 1771
- Nishimoto, U., 1801
- Nitche C. C., 1737
- *Nitzschke, O., 1743
- *Noack, E., 1743
- *Noftzger, C. F., 1907
- Nofter C. R., 1610
- Norbury, A. L., 1761
- *Nordenson, T. E. A.,
1712
- Norman A. G., 1848
- Norris J. H., 1887
- Norris R. J., 1738
- *Norske Aktieselskab
for Elektrokemisk
Industri 1711
- *North C. E., 1922
- *Northrup E. F.
1744
- Nyfeldt, A., 1899
- Nyrid, W., 1949
- Nyren, T., 1863
- Oadkale, U. O., 1823
- *Ober G. & Sons Co.,
1913
- Oberhummer W.,
1877
- Obermann, L., 1826
- Obermeyer, O., 2007
- Oberwender J. L., 1761
- Obraztsov G. D., 1883
- O Daniel H., 1764
- Oe, S., 1801
- *Oeder, G., 1926
- Okkera H. A., 1849
- Oerkov, J., 1867
- Oesterlin M., 1819
- *Österreichisch Ameri-
kanische Magnesia,
A-G., 1966
- *Ofenhaus Ges. m.
b. H., 1790
- Ogata, A., 1919, 1861
- Oguri S., 1964, 2016
- Ohl, F., 1949
- Ohtomo, S., 1871,
1877
- Okahara, K., 1872
- Okamoto K., 1739
- Oldright, C. L., 1776
- Oliver, N., 1921
- *Oliver, H., 1957
- Olson T., 1919
- O'Neill C. D., 1742
- O'Neill, H., 1779, 1780
- Oppel W. W., 1853,
1859
- *Oppr A. G., 1974
- *Oranienberger Chem-
ische Fabrik A. G.,
2006
- *Orchard J. L., 1711
- Orlov, N. A., 1972
- Orthner, L., 1801
- Osborne W. A., 1909
- Ostwald W., 1729
- Orup S., 1726, 1934
- *Otto C., 1951
- *Otto, C. & Co.,
G. m. b. H., 1713,
1976
- Overbeck, F., 1920
- *Overmeyer, O., 2007
- *Owen B. J., 1732
- Owen W. L., 1834
- Owens, E., 1879,
1904
- Ozm T. M., 1753
- Paspe, W., 1916
- *Pacz A., 1791
- Pach, H. O., 1870
- Page, I. H., 1905
- Paiss H. H., 1730
- Paiss, H. S., 1920
- Palache C., 1762
1767, 1768, 1770
- *Palazzo F. C., 1937
- Palmer R., 1858
- *Palmer R. C., 2012
- *Pan L. C., 1743
- Pandera, M., 1831
- Pandara, B., 1848
- Pannach P., 1723
- Papish J., 1757, 1762
- Pardon H., 1802
- Park, B., 1740
- Parla L. R., 2000
- *Parodi DeGinn L.,
1999
- Parrott, P. J., 1911
- Parsons L. B., 1913
- Parsons T. R., 1841
- Parola S., 1890
- Parshin V., 1704
- Passco, P. H., 1770,
1774, 1953, 1964
- *Patent - Treuhänder
Ges. für elektrische
Glühlampen, 1710
- Paternò E., 1801
- Pataardhan G. B.,
2015
- Patwardhan V. G.,
1944
- Paul J. R., 1802
- Pauli, W., 1723
- Paulson M., 1909
- Pavelka, P., 1758
- Pawick, H., 1740
- *Pavle, R., 1789
- Pearce, G. W., 1942
- Pease R. N., 1726
- *Pechin, L. C. F.,
2008, 2013
- *Peder, J. E., 1994
- *Pebbles, C., 1996
- *Pemer K. E., 1962
- Peure F. T., 1874
- Pelanton, H., 1753
- Peola, F., 1901
- *Pena L., 1945
- Perlmann, S., 1877
- Perrot E., 1909
- *Perry J. A., 1975
- Peschke, W., 1910
- Peschel, L., 1787
- Peters, K., 1737
- Peterson B. H., 1721
- Peterson W. H., 1868
- Peutz D., 1865
- Petrashchek W., 1773
- Phanhuoch E., 1822
- Phelice, G., 1890, 1889
- Phelzer, G. J., 1799
- Phelzer H., 1916
- Phelzer, F., 1749
- Pfaff G., 1809
- *Pfeiderer, O., 1743
- *Philippe, F., 1907
- Phillips C., 2872
- Phillips A., 1779, 1786
- *Phillips Engineering
Co. Ltd., 1712
- Phillips P. C., 1784
- Phillips J. B., 2010
- Phillips M., 2019
- *Pick, H., 1926
- *Pieton N., 1996
- *Pier M., 1844
- Pierre R. L., 1918
- Pieskow W., 1879
- Piettre M., 1890
- Pigulevskii V. V., 1979
- *Pial, S., 1954
- Picher, C., 1911
- *Pikington Bros
Ltd., 1963
- Pillet (Mme.), 1870
- Pinal, H., 1761
- Pincussen, L., 1853,
1859
- Pingault, P., 1727
- Pinto, S., 1857
- *Pintach J. A-G,
1973
- Piotrowski, W. J., 1791
- *Pirard O., 2022
- Piron, C., 1776
- Pisumay, M. M.,
1901
- Pittsburgh Equitable
Meter Co., 1970
- Piwowsky F., 1777
- Pizzorno L. V., 1857
- Planck M., 1729
- Plant R., 1720
- Platonov, G., 1893
- Plotnikov, J., 1737
- *P. M. G. Metal Trust
Ltd., 1793
- *Pneumatic Convey-
ance & Extraction
(1929), Ltd., 1709
1711
- *Poetsch W., 1791
- *Poellier G. m. b. H.,
1976
- Pogodin S. A., 1783
- Pokrovskaya, M., 1868
- Polanyi, M., 1716,
1726
- *Poldbüttel, 1792
- Pohard A., 1855
- Pollard C. B., 1817
- *Pomona Werke G. m.
b. H., 1923
- Posthus P., 1816
- Postillon C., 1872
- *Postius O., 1711
- *Porchet, C. C. T.,
1957
- Porter, B. A., 1941
- Porter, M. W., 1719
- Pouchenrieder, H.,
1934
- *Poussin A., 1790
- Potick D., 1856
- Potan, L., 1785
- *Pott A., 1976
- Pourbaix, Y., 1877,
1904
- Pozorinskaya S. A.,
1876
- Powell R., 1781
- Powell, S. G., 1802,
1820
- Prades (Mlle.), 1948
- Pravdyuk, N., 1784
- Pressacker, H., 1821,
1826
- Prentice, F., 1970
- Preobrazhenskiy, A.
M., 1902
- *Preparation indus-
trielle des combus-
tibles 1711,

- Pretschner, F. 1748
 Prévost, C. 1798
 Prévot, J. 1760
 Price, D. J. 1911
 Pringsheim, H. 1900
 Prior, K. 1778
 *Pritchebt & Goll & P. S. Co., Ltd. 1712
 Proske, H. 1945
 *Prückner, R. 1907
 Puche, J. 1886
 Pugh, W. 1750
 Pugsley, L. 1930
 Pulver, H. D. 1778
 Purdy, W. C. 1929, 1930
 Puri, A. N. 1932, 1933
 *Pyrrae Co. Ltd. 1918
 Quattri, O. 1711
 Quaintance, A. 1911
 Quastel, J. H. 1969
 Quenel, P. 1767
 Qureshi, M. 1737
 Raab, W. 1887
 Rabanowicz, J. 1760, 1915
 Rabinowitch, I. M. 1909
 Radef, T. 1886
 *Raffinerie Tielemont, Soc. anon. 2017
 Raiford, L. C. 1915
 Raikbstein, H. A. 1912
 Ramaden, W. 1847
 Randall, M. 1722
 *Randall, W. F. 1795
 Raadon, L. 1875, 1876
 Ranganathas, S. 1803
 Rao, M. R. S. 1721
 Rapoport, J. H. 1961
 Rassow, D. 1741
 Raszeja, P. 1905
 Ratbery, F. 1904, 1905
 Raueh, A. 1747
 *Rauth, F. W. 1921
 *Raw, G. 1711
 Rawdon, H. S. 1786
 Ray, F. F. 1803
 Ray, F. R. 1730
 *Raybaud, L. M. 1922
 Raymond, H. 1708
 Raymond - Hamet, 1902, 1903, 1909
 Razuvaev, G. A. 1831
 Read, M. T. 1714
 Rebagnati, E. E. 1762
 Re, P. M. 1856
 *Rector Gausler Co. 1973
 Redlich, C. 1961
 Reed, G. B. 1869
 Reed, H. M. 1920
 Rees, W. J. 1960
 *Reeset, H. 1958
 Regel, O. 1927
 Regaer, W. 1828
 Régner, J. 1907
 Reihlen, H. 1864
 Reilly, J. 1828
 *Rehmann, E. 1791
 Reimer, W. 1931
 Reia, C. R. 1809
 Reiner, M. 1898
 Remon, Neulert 1729
 Remy, T. 1919
 Reosaudie 1795
 Reakin, W. G. 1973
 *Reukte, B. 1965
 Rewahl, H. 1915
 *Rewahl, B. 2019
 Rewuridge, A. t. 1901
 Reynolds, A. F. 1918
 Reynolds, N. D. 1736
 Rhind, D. 1917
 Rhoads, C. P. 1900
 Rhoad, G. 1720
 Ribotti, Lwonne, G. 1900
 Rice, O. K. 1738
 Richard, D. W., Jr. 1883
 *Richards, F. M. 1972
 Richards, F. T. 1777
 *Richards, L. C. 1922
 Richardson, C. H. 1911
 *Richardson, C. M. 2021
 Richardson, H. H. 1910
 Richardson, R. 1800
 *Richfield Oil Co. of California, 1983
 Richter, G. H. 1861
 Ridgway, J. H. 1919
 *Riech, F. 1967
 Riechmeier, O. 1726
 *Rietel, J. D. E. de Hafe A.-G. 1922
 van Riemstyk, M. 1836
 Righellato, B. C. 1720
 *Riley, T. N. 1715
 Rimbach, R. 1774
 *Ringer, F. 2000
 Riokenbach, W. H. 1908
 Rittenhouw, C. G. 1909
 Rivoler, R. 1899
 Roark, R. C. 1940
 Robert, 1929
 Roberti, G. 1969
 *Roberts, E. G. L. 1712
 Roberts, H. N. 1926
 *Robertson, I. W. 2022
 Robertson, M. E. 2018
 *Robinson, E. V. 1709
 Robltschek, J. 1961
 Roche, B. de la 1756
 Römer, G. H. 1720, 1828
 Roes, H. F. 1856, 1886
 *Rogers, F. M. 1983, 1944
 Rogers, J. N. 1731
 Rogers, W. R. 1770
 Rohmann, A. 1721
 Rolle, R. T. 1761
 Roll, P. 1740
 Rolleson, G. K. 1736
 Romanoff, A. L. 1888
 Roman, G. 1761
 Romig, I. W. 1959
 Rompe, H. 1735
 Rona, P. 1818
 Rondelli, H. 1888
 Rondoni, P. 1875
 Rose, A. R. 1865
 Roser, E. 1973
 Rosellat, J. I. 1912
 Rosenbaum, B. 1977
 Rosenberg, N. J. 1778
 Rosenthal, D. 1786
 Rosenthal, S. M. 1919, 1908
 Rosenthal, W. 1788
 Rosenthaler, I. 1871
 *Rosling, P. 1745
 Ross, J. 1802, 1803
 Ross, W. A. 1911
 Rosson, F. D. 1727
 *Roth, H. W. 1947
 Rothman, A. 1907
 Rothman, T. 1903
 *Rouni, A. A. 2020
 Rowe, R. C. 1921
 Rowlette, A. P. 1859
 *Ruben, S. 1710
 Rucker, M. P. 1911
 Rudnew, G. T. 1861
 Rudolfs, W. 1928, 1929
 *Rüppel, W. 1711
 Rüst, H. 1729
 *Rütgerswerke, A. G. 1842
 Rupe, H. 1820
 *Russ, F. F. 1741
 Russell, H. E. 1917
 *Rutenberg, A. 1866
 *Ruth, G. Akt.-Ges. 2011
 Ryshukh, P. 1979
 Rydman, M. 1876
 Rzymkowski, J. 1717
 Sabatay, S. 1817
 *Sachs, A. P. 1985
 Sabat, A. B. 1900
 *Sachse, H. 1830
 *Sachtleben, A. G. für Bergbau und Chemische Industrie 2011
 Sack, H. 1739
 *Sailer, G. H. 1974
 *Sale-T. Stat. Co. 1709
 Nagar, J. 1913
 Nagorshew, B. 1750
 Nahai, P. N. 1803
 Nahshi, Y. 1876
 Schavatschke, D. L. 1911
 *S. A. I. G. (Soc. anon. international Guadagn), 1959, 1965
 *Sailer, W. 2019
 St. v. Nary Szabó 2017
 Sakurai, J. 2018
 Salmaag, H. 1961
 Salmon, W. D. 1878
 Salmony, Karsten, A. 2001
 Salto, L. 1827
 Samola, N. A. 1731
 Sanchez, J. A. 1762
 Sandegren, R. 1774
 Sanders, R. 1908
 Saeford, H. N. 1878
 Sanharan, G. 1888
 Sano, Y. 1931
 Santmyers, R. M. 1777
 Sarasin, J. 1857
 Sarkar, P. B. 1731
 Satri, G. N. 1878
 Sato, M. 2016
 Sattler, G. 1926
 Sauer, F. 1733
 Sauer, J. 1859
 Sauer, O. K. 1915
 Sauerwall, I. 1760
 *Sauters, K. H. 1711
 *Savage, J. 2016
 Savell, W. I. 2000
 Sax, P. 1907
 Scales, P. M. 1917
 Schacht, G. 1725
 Schaud, R. E. 1706
 Schaller, W. T. 1760
 Schanbacher, K. L. 2000
 Schatler, R. 1770
 Scham, K. 1718
 Scheffer, F. 1938
 *Scheffer, W. 1921
 *Schibler, J. 2011
 Schenk, P. W. 1740
 Schenk, R. 1883
 Scheps, M. 1891, 1908
 Scherer, P. C. Jr. 1900
 *Schering - Kahlbaum A.-G. 1931, 1991
 Scherings, K. 1803
 Schey, S. 1730
 Schleich, M. 1881
 Schiele, J. 1725
 Schiff, F. 1907
 *Schiffner, H. J. 1701
 *Schimmel & Co. A. G. 1830
 Schmale, R. 1801
 Schirmer, M. 1887
 Schivo, A. J. 1887
 Schilder, P. 1726
 *Schlegel, W. 1712
 Schlecker, E. 2016
 Schlossmacher, K. 1719
 *Schlotterhose, C. 2018

- Stewart, C. P., 1932, 1953
 Stewart, E. D., 1920
 Stewart, F. G., 1972
 Stewart, H. J., 1883
 Stewart, J. R., 2009
 Stewart, T. D., 1903
 Stewart, W. W., 1799
 Stich, C., 1754
 *Stich, F., 1945
 Stieger, G., 1907
 Stiepel, C., 2018
 Stillson, C. B., 1762
 Stutz, J., 1920
 *Stokes, R. A., 1712
 *Stokes, W. P., 1972
 Stoll, L., 1969
 Stone, I., 1707, 1759, 1762
 *de Stordeur, R. J. A., 1976
 Storck, K. H., 1721
 *Story, H. E., 1712
 Stotz, R., 1778
 Strasburger, J., 1849
 *Strauch, C. B., 1950
 Strauss, H., 1894
 Strauss, H. A., 1911
 Strauss, M. L., 1853
 Strauss, S. G., 1907
 Streeter, L. R., 1942
 *Stressau, R., 1743
 Striebel, H., 1715
 *Strindlund, J., 1996
 *Strobaach, O., 1923
 Strom, D. A., 1979
 *Stuart, J. F. W., 1962, 1963
 Stuart, M. C., 1923
 Stuckey, J. L., 1773
 Sturges, W. S., 1912
 Stutz, G. F. A., 1735
 Styrikovich, V. L., 1901
 Süss, O., 1833
 *Sugimura, I., 1790
 Sullivan, J. D., 1775
 *Sullivan, J. H., 1963
 Sulman, H. L., 1774
 Sumu, M., 1875
 Sundelin, G., 1921
 Sundius, N., 1764
 Süßer, A. P., 1886
 Suponitskii, L. V., 1953
 *Suzerik, L., 1839
 Suter, C. M., 1816
 Suzuki, B., 1800, 1801, 1836
 Svecnikov, V.-G., 1782
 *Svenska Akkumulator Aktiebolaget Jungner, 1709
 Swallow, H. T. S., 1960
 *Swift & Co., 1923
 Swingle, M. C., 1913
 Swope, H. G., 1929
 Swope, W. D., 1917
 Sym, R. A., 1849
 Szécs, B., 1778
 Szucs, F., 1864
 Szécs, S., 1945
 Taber, S., 1769
 Tailandier, M., 1963
 Takagi, T., 1870
 Takahashi, T., 1802
 Takahashi, T., 1877
 Takayama, Y., 1913
 *Takemura, K., 2011
 Tamm, I., 1729
 Tammann, G., 1724, 1726, 1784
 *Tams, J. E., 1964
 Tanaka, M., 1953
 Tanaka, S., 1717
 Tankó, B., 1859
 Tapadinhas, J., 1850
 *Tartas, L., 1713
 Tatarski, H., 1902
 Tausz, J., 1760
 Tavaris, C., 1946
 *Taylor, A. O., 1789
 Taylor, E. R., 1781
 Taylor, F., 1741
 Taylor, H. S., 1737
 *Taylor, J. H., 2011
 *Taylor Wharton Iron & Steel Co., 1792
 Tchakurian, A., 1751
 *Technical Research Works, Ltd., 1710
 Telesnaw, K., 1876
 Temple, R. L., 1784
 Teorell, T., 1754
 Terrance, E. F., 1864, 1881
 *Terzan, H. G., 1976
 Thayer, A. C., 1771
 Thes, E. R., 2013
 Theriault, E. J., 1929
 Thielke, R. C., 1823
 *Thies, K., 2004
 Thilo, E., 1950
 Thivolle, L., 1884
 *Thomas & Hochwalt Laboratories, Inc., 1973
 *Thomas, J., 1843, 2002, 2004
 Thomas, J. S., 1750
 *Thomas, L., 1709
 Thomas, P. E., 1857
 Thomas, W. H., 1950
 Thompson, H. B., 1950
 Thompson, J. G., 1783
 Thompson, J. T., 1899
 *Thompson, L. F., 1713
 Thompson, F. K., 1908
 Thompson, T. C., 1714
 Thompson, W. C., 1830, 1903
 *Thomsen, J. E., 1953
 Thomson, D. F., 1909
 Thomson, J. R., 1942
 *Thomson, K. F., 2002
 Thorne, B., 1909
 Thurber, F. H., 1823
 Thwing, C. F., 1714
 *Tibbenham, L. J., 1794
 Tiemann, F., 1912
 *Tiquant, H. T., 1793
 Tinker, J. S., 1864
 Tipson, R. S., 1805
 Tisdale, E. S., 1930
 *Titanium Pigment Co., Inc., 2011
 Tittler, R. P., 1867
 *Toback, I., 1948
 *Toda, A., 1994
 Toennies, G., 1802
 Tomarchio, G., 1964
 *Tomlinson, J. N., 1712
 *Tomlinson (Rockdale), Ltd., 1712
 Tommasi, G., 1936
 Tomckwaere Debergh (Mme), 1953
 Tomgrutti, M., 1979
 *Tomkret Ges. m. b. H., 1966
 Torosyan, G., 1977
 Torberg Jensen, S., 1939
 Towne, A. P., 1775
 *Tramer, L., 1967
 Trask, P. D., 1772
 *Trautwein, H., 1974
 Trautz, O. R., 1763
 Travers, A., 1726
 Trestraal, G., 1771
 Trey, F., 1717
 Traka, J., 1817
 Trumpp, D., 1735
 Truninger, E., 1937
 Tschertkow, L., 1997
 Tschinkel, H., 1802, 1873
 Tschirch, E., 1760
 Tso, E., 1875
 Tucca, P., 1786
 Tuckfield, F., 1853
 Tüchel, S., 1762
 *Turkington, V. H., 2012
 *Turk, D. F., 2020
 Tyler, P. M., 1777
 Udaondo, C. B., 1909
 Uhlenhuth, F., 1869
 Uhlmann, P. W., 1969
 Ullrich, A., 1810
 Umstätter, H., 1959
 *Unilever, N. V., 1922
 Uno, D., 1784
 Urazov, G. G., 1785
 Urbach, P., 1906
 Urbach, F., 1747
 Urbain, A., 1895
 *Urbach, O., 1922
 Urmonston, J. W., 1737
 *Urtum, V., 1994
 Uyeda, T., 1778
 Vadara, E., 1785
 Vahl, L., 1720
 Vahlquist, B., 1859
 Valdayanathan, V. I., 1732
 Valda, O., 1923
 *Vale, E. C. H., 1996
 Valette, G., 1903
 Van Cleave, R., 1714
 *Vanderbilt, R. T. Co., Inc., 2016
 Van Dyk, J. C., 1915
 Van Hoon, F. R., 1770
 Vardavas, S., 1773
 Vaskariy, H., 1899
 Vashadri, P., 1893
 Vasilou, V. A., 1864
 Vaudin, L., 1887
 Vavou, G., 1909
 Vecchiotti, L., 1821
 Vegerel, G., 1944
 van de Velde, J., 1902
 Vekovskii, A., 1978
 Velje, L., 1722, 1861
 Venn-Dantova, E., 1911
 *Vereinigte Alumin-Jum Werke, A.-G., 1799
 *Vereinigte Glühlampen und Elektrische Werke, A. G., 1710
 *Vereinigte Huttenwerke Bloch & Hirsch, 2019
 *Verbeke, T. H., 1945
 Vermaad, V. I., 1873
 Verne, J., 1851
 *Verpla, P., 1945
 Veron, D., 1922
 Versluis, J., 1773
 Verraz, P., 1899
 Vessal, G., 1717
 Viale, G., 1858
 *Vickers-Armstrongs, Ltd., 1793
 *Vickery, P. W., 1996
 *Victor Chemical Works, 1934
 *Victor X-Ray Corp., 1709
 *Victoria Rubber Co., Ltd., 2021
 Vieira, O. de S., 1918
 *Vilain, C., 1942
 Villars, M., 1904
 Villars, D. S., 1731
 Virgilio, B., 1863
 Visser, G. H., 1735
 Viterba, E., 1747
 Voegtlin, C., 1903
 v. Vogel, H. U., 1751
 *Vogt, C. W., 1922
 Voigt, B., 1717
 Voichenko, U. M., 1933
 Volkmann, W., 1707
 *Vollmann, H., 2006
 *Vondran, A., 1925
 Voronov, A. I., 1982
 Vos, H., 1893, 1894
 Voss, W., 1797, 1798
 Voznesenskii, S., 1714
 Wachsmuth, E., 1778
 Wada, S., 1884
 *Wade, H. C., 1838
 Wacker, B., 1936
 Wagner, A., 1781
 Wagner, C., 1724

- Wagner, E. 1748
 *Wagner, G. R., 1956
 Wagner, H. 1943
 Wagner O. H., 1737, 1971
 Wagner, R. D., 1943
 Wakeman, A. M., 1892
 Waldo, J. H., 1821
 *Waldron, P. H., 1963
 Walker, A. W., 1896
 Walker, D. S., 1848
 Walker, V. P., 1853
 *Walker, W. 1791
 *Walker, R. S., 1712
 Wallace, C. H., 1900
 *Wallace, W. S. 1789
 Waller, I., 1729
 *Wallon, O. H., 1973
 Walster, H. L., 1932
 Walter, C. M., 1970
 Walters, F. M., Jr. 1783
 Walters, L. S., 1756
 Wallis, A., 1818
 Walton, R. P., 1949
 Warkony, J., 1877
 Wasmbeck, P., 1803
 Warner, W. B., 1783
 Warren, D. B., 1764, 1785, 1786
 Warren, S. L., 1911
 *Wernick, B. O., 1937
 *Wernick, M., 1937
 Wassermann, C., 1717, 1784
 Waterman, H. I., 1729
 Watson, C. H., 1708
 Watts, H. P., 1909
 *Watzinger, G., 1906
 Weitenauer, A., 1772
 Waugh, T. R., 1802
 *Webber, C. S., 1749
 Weber, F., 1768, 1936
 Weber, K., 1746
 Weber, W., 1877, 1886
 Webers, R., 1953
 Webster, J. E., 1874
 *Weck, G., 1703
 Weed, A., 1940
 *Wehe, H. D., 2010
 *Weiland, H. J., 1821
 *Werner, R., 1740
 *Weise, H., 1971
 *Weiser, H. B., 1722
 *Weiser, I., 1881
 *Weiskopf, P., 1939
 *Weiss, P., 1896
 *Weinbach, K., 1811, 1813
 *Weissberg, K., 1760
 *Weitzverkhaskabel, G. m. b. H., 1743
 *Wallecoma Foundation Ltd. 1931
 *Wells, C., 1783
 *Wells, R. C., 1763
 *Welsh, J. N., 1827
 *West, S., 1830
 *Werder, J., 1945
 *Werner, K., 1903
 *Wernicke, R., 1801, 1896
 *Werschen - Weiden Jelier Braunkohlen A.-G., 1908
 *van Wert, L. R., 1779
 *Wertheim, R., 1812, 1838
 *Wertz, W., 1853
 *West, J., 1766
 *Westall, L., 1806
 *Wever, F., 1718
 *Weyman, J. H., 1713
 *Wheatley, R., 2021
 *Wheeler, P., 1960
 *Wheeler, T. S., 1853
 *Whipple, G. M., 1912
 *White, A. H., 1870
 *White, C. B., 1953
 *Whitehead, A. N., 1739
 *Whitehead, T. H., 1782
 *Whitehead, W., 1994, 2007
 *Whitfield, R., 1783
 *Whitmore, W. P., 1737
 *Whitutt, M. L., 1876
 *Wiberg, B., 1751
 *Wichard, A., 1927
 *Wickmann, F. A., 1923
 *Widemann, R. V., 1961, 1962
 *Widmark, O., 1869
 *Wiegand, H., 1916
 *Wiegand, W. B., 2023
 *Wiegner, B., 1997
 *Wien, M., 1735
 *Wieruchowski, M., 1879, 1880, 1904
 *van Wijk, W. R., 1735
 *Wilhoof, H. L., 1817, 1909
 *Wilhelm, R., 1976
 *Willard, C. T., 1791
 *Willson, C. O., 1940
 *Willstätter, R., 1921
 *Willer, B., 1773
 *Wilson, W. C., 1791
 *Wimmer, J., 1843
 *Winkler, J., 1791
 *Winnograd, H., 1742
 *Winter, H., 1808
 *Winter, P., 1948
 *Wintersta, R., 1833
 *Witkenus, H., 1963
 *Witlich, L., 1907
 *Witold, H., 1897
 *Witka, P., 2015
 *Witold, 2015
 *Wode, G., 1918
 *Wohlegemuth, J., 1852
 *Wolbach, G., 1851
 *Wolf, H., 2014
 *Wolf, L., 1741
 *Wolf, L., 1712
 *Wolfe, H. C., 1780
 *Wolfriden, J. H., 1728
 *Wolf, H., 1760, 2015
 *Wolfram, A., 1845
 *Wolfsch, G., 1733
 *Wolman, A., 1930
 *Woo, S. C., 1725
 *Woodell - Duckham (1929), Ltd., 1925
 *Woodbridge, J. L., 1743
 *Woodford, A. O., 1784
 *Woodroffe, P. K., 1974
 *Woog, P., 1923
 *Woolf, J. A., 1775
 *Wright, L., 1741
 *Wright, F. C., 1769
 *Wright, R. A., 1922
 *Wright, S. L. Jr., 1892
 *Wuhr, C., 1963
 *Wyckoff, R. W. G., 1719
 *Wythes, V., 1793
 *Yamaji, K., 1943
 *Yamamoto, T., 1718
 *Yamanouchi, T., 1861
 *Yamasaki, M., 1873
 *Yang, C. S., 1903
 *Yentir, W., 1988
 *Yemitsky, N. N., 1901
 *Yamamoto, M., 1790
 *Yatlov, V. S., 1951
 *Ylla Conte, J., 1956
 *Yokoyama, Y., 1836
 *Yoshida, R., 1801
 *Yoshimaru, Y., 1878
 *Young, A. C., 1770
 *Young, P. O., 1796
 *Young, H. B., 1978
 *Young, J. W., 1770
 *Youngburg, G. H., 1857, 1899
 *Youngburg, M. V., 1857, 1899
 *Zach, C., 1945
 *Zabaria, A., 1977
 *Zahradnický, J., 1733
 *Zaitchuk, A., 1841
 *Zemouruev, G. M., 1783
 *Zerbstmeyer, O., 1986
 *Zechmeister, L., 1817
 *Zechner, L., 1946
 *Zelner, H. P., 1934
 *Zerger, H., 1829
 *Zerkert, O., 1948
 *von Zelewski, R., 1791
 *Zelnick, N. D., 1963
 *Zemlyanskiy, V. P., 1771
 *Zenneck, J., 1729
 *Zerbe, C., 1968
 *Zhuze, V., 1978
 *Ziegler, K., 1829
 *Zimmerley, S. R., 1707
 *Zipperer, L., 1969
 *Zlataroff, A., 1891
 *Zondek, B., 1853
 *Zachack, F. H., 1959
 *Zublin, E. W., 1982

CHEMICAL ABSTRACTS

Editor: E. J. Crane, Ohio State University, Columbus, Ohio.
Associate Editors: Elmer Hockett, Helen Gama and Leonard T. Capall
Assistant Editors

Al S. Anderson	D. J. Demarest	F. B. Hawk	L. A. Olney	J. J. Skinner
W. L. Badger	Alden H. Emery	P. H. Hines	A. M. Patterson	W. R. Siemen
G. R. Barton	W. H. Himes	C. H. Karr	T. G. Phillips	C. G. Storm
Edward Bartow	W. P. Paraghar	H. A. Lepper	A. N. Richards	R. M. Symmes
S. R. Benedict	A. C. Peetles	H. K. Marshall Jr.	Richard Rimbach	P. F. Tindeshill
P. C. Black	C. G. Pink	A. R. Middleton	Allen Rogers	H. G. Wells
P. L. Brown	C. N. Frey	Harlan S. Miner	C. A. Roubier	C. J. West
C. F. Curren	H. W. Lullitt	Chas. B. Morrey	A. H. Sabin	R. T. Wherry
J. K. Dale	R. A. Cooper	C. H. Moore	I. P. Schairer	E. P. Wightman
C. C. Davis	W. T. Hall	W. A. Noyes Jr.	R. Scherbel	J. C. Witt

CONTENTS

Patents .. .	1	44. The Fermentation Industries .. .	2337
Author Index .. .	ii	45. Pharmaceutical Chemistry .. .	2442
1. Apparatus and Plant Equipment .. .	2024	46. Acids, Alkalies, Salts and Sun-	2447
2. General and Physical Chemistry .. .	2030	dries .. .	2447
3. Subatomic Phenomena and Radio-		47. Glass, Clay Products, Refractories	2454
chemistry .. .	2045	and Enamelled Metals .. .	2454
4. Electrochemistry .. .	2044	48. Cement and Other Building Ma-	2443
5. Photography .. .	2044	terials .. .	2443
6. Inorganic Chemistry .. .	2044	49. Fuels, Gas Tar and Coke .. .	2444
7. Analytical Chemistry .. .	2070	50. Petroleum, Lubricants, Asphalt	2472
8. Mineralogical and Geological		and Wood Products .. .	2472
Chemistry .. .	2072	51. Cellulose and Paper .. .	2482
9. Metallurgy and Metallography .. .	2041	52. Explosives and Explosions .. .	2484
10. Organic Chemistry .. .	2151	53. Dyes and Textile Chemistry .. .	2485
11. Biological Chemistry .. .	2147	54. Paints, Varnishes and Resins .. .	2497
12. Foods .. .	2104	55. Fats, Fatty Oils, Waxes and Soaps .. .	2414
13. General Industrial Chemistry .. .	2410	56. Sugar, Starch and Gums .. .	2412
14. Water, Sewage and Sanitation .. .	2416	57. Leather and Glue .. .	2421
15. Soils, Fertilisers and Agriculture)		58. Rubber and Allied Substances .. .	2442
Poisons .. .	2422		

PATENTS

A statement giving information as to how to obtain patent specifications and drawings, both United States and foreign, is to be found on page i of *Chemical Abstracts*, 24, No. 5

Published by THE AMERICAN CHEMICAL SOCIETY, which also publishes *The Journal of the American Chemical Society* and *Industrial and Engineering Chemistry*.

Communications for the editor should be addressed to E. J. Crane, Ohio State University, Columbus, Ohio.

Remittances for annual dues (\$15.00) in the American Chemical Society or for subscriptions to this *Journal* (\$7.50 per year), changes of address, orders for subscriptions or back numbers, claims for raising numbers and all matters intended for the *Proceedings* should be sent to the Secretary, Charles L. Parsons, Mills Bldg., Washington, D. C. Foreign postage is extra, being \$4.50 per year added to the dues of members residing outside the United States, its possessions and the countries of the Pan American Union and \$1.50 per year for *Chemical Abstracts* alone, except that only one-half of these amounts is added for journals going to Canada. The countries of the Pan-American Union are Argentina, Bolivia, Brazil, Chile, Colombia, Costa Rica, Cuba, Dominican Republic, Ecuador, Guatemala, Mexico, Nicaragua, Paraguay, Peru, Republic of Honduras, Republic of Panama, Salvador, Spain and Uruguay.

The Council has voted that no claims will be allowed for copies of journals lost in the mails unless such claims are received within sixty days of the date of issue, and no claims will be allowed for issues lost as a result of insufficient notice of change of address. "Missing from files" cannot be accepted as the reason for honoring a claim.—Charles L. Parsons, Secretary

AUTHOR INDEX

[An asterisk (*) indicates that the reference is to an abstract of a patent]

- Aartema, G. A., 2053
 Abadi G., 2245
 *van Ackere, J., 2275
 Ackermann, D., 2159
 Adams, J. M., 2035
 *Adams, L. V., 2310
 Adams, R., 2135
 Adams, W. W., 2292
 Adam, M., 2167, 2184
 Adcock, F., 2054
 Adcock, P., 2115
 Agazziotti, A., 2150
 Agnew R., 2151
 Agre, C., 2036, 2091
 *Aitchison, J. H., 2111
 Aikman, C., 2154
 *Akrynia, L. S., 2294
 *Alt-Ges. Brown
 Boveri & Co., 2060
 *Alt-Ges. der Mas-
 schinenfabriken
 Escher Wym & Co.,
 2251
 *A.-G. für Special-
 bauteile, 2215
 *Altendickelbet Niro,
 2025
 *Alberts, W., 2105
 Albrecht, W. H., 2032
 Albu, H. W., 2309
 Alexander, H. H.,
 2054
 Alexander, L. T., 2229
 Alexov, A. L., 2102
 Alfthan, K., 2071
 Altoparo, R., 2246
 Alton, A., 2265
 *Alvard, H., 2155
 Alder, J. F., 2095,
 2099
 Alder, W. W., 2169
 Akner, J., 2182, 2008
 Alfred, M., 2245
 Albertson, H., 2036
 *Alumina - Indus-
 trie, A.-G., 2103,
 2215
 *Alumina Werke, A.-
 G. Rorschach, 2111
 Amelink, J. F. H.,
 2241
 *Americas Rotig
 MCI Co., 2110
 Amoson, R., 2160
 Amster - Sardo, J.,
 2143
 Anderson, A. C., 2204
 *Anderson, C. O.,
 2059
 Anderson C. S., 2055
 *Anderson, E., 2062
 *Anderson P., 2272
 Ando, N., 2045
 Andrichew, M.,
 2227
 Andrew, J. H., 2053
 *Andrew, D. S., 2272
 Andrew, R. S., 2265
 *Anglo-Americas O.,
 Co., Ltd., 2251
 Anthony, A. J., 2024
 Anyama, S., 2150
 *Aprar, P. A., 2240
 *Appareils et Evapor-
 teurs Kestner, 2251
 Appleyard, R., 2045
 Aral, H., 2074
 Arbatsky, J. W., 2269
 Archbutt, S. L., 2100,
 2101, 2102
 Archibald, R. M.,
 2033
 *Arend, A., 2110
 *van Arkel, A. H.,
 2061
 *Armstrong A., 2264
 *Armstrong, J. J.,
 2221
 Armstrong, T. N.,
 2045
 Arnaud, G., 2234
 Arndt, F., 2145
 Arnot, F. L., 2046
 *Arnott, J. B. E. E.,
 2029
 Arshinov V. V., 2079
 Arthur, R., 2159
 Arup, P., 2006
 Ashida, Y., 2145
 Asnosa, C., 2157
 Asher, L., 2179
 Asada Y., 2190
 van Asperen de Boer,
 S. R., 2165
 *Associated Electrical
 Industries, Ltd.,
 2029
 Astbury W. T., 2297
 Astruc, H., 2237
 Atanasiu, I. A., 2277
 *Aubert & Duval
 Frères, 2109
 Aubert, H., 2155
 *Auer, H., 2292
 *Auer, L., 2153, 2279
 2311, 2312
 Auerbach, M., 2227
 August, A., 2008
 *Autogen Camerama
 Inter, A.-G., 2274
 Avest, A. G., 2314
 Avery, W. & T.,
 Ltd., 2239
 Avetisova, A. N.,
 2160
 Avtomotova, E., 2160
 Ayv, A., 2112
 *Litz Research Lab-
 oratories, Inc., 2250
 *Azoprene (Società
 anon. per la fabbri-
 cazione dell'am-
 moniac anidro e
 prodotti derivati)
 2250
 Baas-Becking G. M.,
 2172
 Babcock, S., 2250
 Balbo, A. E., 2074
 Backus, H. A., 2214
 *Bader, W., 2155
 *Baer, J., 2332
 Bag, A., 2317
 Bagchi, R., 2252
 Baker, A., 2095
 *Bailey, H. S., 2106
 *Bailey, J., 2259
 Baker, W. L., 2277
 *Bakke, G. m. b. H.,
 2211
 Baker, H. C., 2042
 Baker, J. C., 2026
 Baker J. W., 2130,
 2121
 Baker R. C., 2024
 Baker W. E., 2255
 Balh, T. N., 2076,
 2077
 *Bakhuizen, J. W.,
 2029
 Balh, R. T., 2221
 Baldoni, A., 2245
 Baldrecca, G., 2225
 *Baldwin, J. T., 2212
 Balke, R., 2227
 *Balz Bros. Co., 2353
 Balz, O., 2141
 Balz, E. C. C., 2220
 Bamberger, E., 2125,
 2126
 *Bancet, P. A., 2029
 Banerjee, K., 2025
 Banet, C., 2145
 Banbour H. G., 2202
 Banerji, J., 2201
 Ban, M. E., 2210
 Banmore C. B., 2265
 Barnea, M. E., 2216
 Barnett, E. de B.,
 2140
 Barnett, R. M., 2309
 *Baron, J. T., 2274
 *Barraet, E. S., 2255
 Barrett, W. H., 2045
 Bartha, W. B., 2055
 Barth, F. W., 2051
 *Barth, H. L., 2111
 Barlow, E., 2215
 Bartelst, F., 2208
 *Bartsch, R., 2290
 *Basler M., 2155
 Bassett H. N., 2215
 Bass, R., 2285
 *Bassett E. M., 2292
 Bates, F. H., 2260
 Batson, R. C., 2056
 *Battenby, J. W.,
 2253
 Basch, M., 2176
 Bauer, O., 2101, 2121
 Baule, E. E., 2195
 Bauman, J. E., 2221
 Baumgarten, P., 2120
 *Baumgarten, F., 2146
 Baxter G. F., 2066
 *Bayer, E. C., 2263
 *Bayerische Stock-
 Werke A.-G., 2250
 Bayle, R. M., 2054
 *Baylis, W. S., 2251
 *Beade, G. W., 2009
 Bean, H. J., 2117
 *Beans, H. T., 2254
 Becker, R., 2153
 Beck, G., 2046
 *Beck, H. P., 2053
 *Beck, K., 2299
 *Becker, A. C., 2079,
 2250
 Becker, H., 2178
 Becker, K., 2236, 2042,
 2091
 *Beckert, F. M., 2106
 Beccarello, J., 2052
 Bedding, W. C., 2319
 Beerman, H., 2054
 *Begrov, I. L., 2310
 Behrman, A. S., 2217
 *Benner, G., 2209
 Belarsky, E., 2224
 *Bell, J. E., 2279
 Bell, J. M., 2051
 Belcher, M., 2079
 *Bellis, A. E., 2106
 Bellob, E., 2185
 *Belsley, J. P., 2157
 Belzoni, G., 2199
 Beljare, N. T., 2093
 Beljare, S. M., 2184
 *Bembridge J. F. A.-G.,
 2290
 *Bendix, R., 2209
 Bentham, E., 2051
 *Benner, R. G., 2061
 Bennett, H. G., 2225
 Bennett, O. G., 2036
 *Bennett, R. A., 2106
 Bennish, A., 2067
 Benso, P., 2191
 Benson, W. L., 2119
 Bent, H. E., 2023
 *Bentley - Himey, A.,
 2026
 Bentkovich, L., 2261
 Benton, A. P., 2039
 *Benton, W. A., 2229
 Benton, W. E., 2026
 *Berdell, T. van D.,
 2235
 Berg, R., 2174
 Berg, R. N., 2155
 Bergam, G., 2181
 *Bergendorfer Eisen-
 werk A.-G. Astra-
 Werke, 2209
 *Bergel, S., 2246
 Bergman, M., 2151,
 2224
 Bergqvist, O., 2036
 Berk, A., 2156
 *van Berkel, C. F. M.,
 2312
 *Berlinerthal, J., 2205
 *Berry, J. A., 2106
 *Beranek, A., 2238,
 2220
 *Berry, A. F., 2259
 *Berry, W. M., 2030
 *Berthmann, W.,
 2220
 Berthoud, R., 2065
 Bertus, C., 2229

- *Beecheynick, W., 2024
 Bestehorn, R., 2269
 Betha, H., 2016
 *Bethel, A. R., 2234
 Bettet, S., 2201
 Beutner, R., 2162
 Bewilogua, L., 2050
 *Beyrich, R., 2249
 Bhagwat, V. K., 2144
 Bhagwat, W. V., 2013
 Bhatnagar, S. S., 2124
 Bhattacharya, A. K., 2032
 Bhattacharya, R., 2206
 Bianchi, C., 2310
 Bickel, A., 2192
 Bickenbach, W., 2176
 Bidwell, H., 2025
 Biffrhy, T., 2201
 Bieglow, H. E., 2126
 Bulmann, R., 2012
 Bull, A. H., 2191
 Bulter, G., 2123
 Bulter, J., 2054
 Buncer, V., 2177
 *Binkley, M. J., 2027
 Nina, A., 2192
 Bird, B. M., 2263
 *Birmingham Alumina
 lum Casting (1903)
 Co., Ltd., 2104
 *Blieckel, J. A., 2274
 *Blieckhoff, F. R., 2215
 Bishop, R. O., 2229
 Bishop, W. B. S., 2174
 Bistrzycki, A., 2142
 *Black, J. A., 2275
 Black, R., 2140
 Blackburn, W. H., 2044
 Blair, C. A., 2225
 Blanchon, H., 2124
 Blankenburg, C., 2127
 Blankensfeld, A., 2161
 *Blaschke, E., 2290
 Blau, R., 2024
 *Blecker, W. F., 2278
 Bloch, R., 2042
 *Block, D. J., 2256
 Blom, A. V., 2214, 2304
 Blum, I., 2265
 Blum, W., 2036
 *Blythen, S., 2213
 Boby, W., 2218
 Bockhammer, 2278
 Boehm, T., 2145
 *Boehringerr, C. F., &
 Soehne, 2210
 *Bogart, G. B., 2029
 Boggs, C. R., 2277
 Bogoyavlenskii, L., 2049
 Bokil, K. V., 2126
 *Bodmann, H., 2227
 Bolm, F., 2242
 Bolton, L. W., 2102
 *Bongrand, J. E. C., 2203
 Bonhoeffer, K. F., 2050
 *Bonnar, T. K., 2254
 Bonnet, R., 2172
 Bonrango, A., 2211
 Booth, N., 2184
 Boelaski, P., 2178
 Boelov, P. P., 2014
 *Borndorf, R., 2060
 Bornstein, A., 2191
 Borra, V., 2203
 *Borrowman, G., 2222
 Boscock, H., 2142
 *Bosse, A., 2250
 Borst, M., 2174
 Bosward, M., 2103
 Bouckaert, J. J., 2202
 Boughton, W. A., 2277
 Bourgin, D. C., 2023
 Bourguignon, G., 2202
 Bouquet, F. W., 2113
 *Boys, A., 2239
 *Boyer, S., 2024, 2156
 Brabant, 2218
 *Bradenhorst Elektro
 maschinen Ges., 2202
 Brackmiller, L., 2182
 *Bradley, A. K., 2274
 Bradd, W. E., 2037
 Brandrup, W., 2241
 Brandt, W., 2126, 2127
 *Brandwood, J., 2204
 Bratting, K., 2043
 *Braun, C. F., 2024
 v. Braun, J., 2124
 Braun, W., 2127
 Braunket, W., 2049
 Braunsdorf, H. A., 2232
 Bravo, G. A., 2292
 Bray, G. W., 2189
 *Brehmer, F. G., 2216
 Brekhaupt, T., 2219
 *Brenn, R., 2223
 Brennan, G. L., 2267
 Bretnschneider, H., 2144
 Bremer, F. H., 2169
 Brewster, R. Q., 2125
 Briley, L. J., 2102
 Brickwedde, F. G., 2023
 *Bridger, T. B., 2109
 *Bridgman, F. W., 2252
 Briggner, R., 2245
 *Brightman, R., 2202
 Brill, R., 2166
 Briner, E., 2127
 *Britlich, Remberg,
 Ltd., 2220
 *British Celanese
 Ltd., 2304, 2305,
 2312, 2313
 *British Hartford Felt
 mill Syndicate,
 Ltd., 2258
 *Britton, E. C., 2153,
 2156
 *Brook, H. J., 2223
 *Broderick, F. W., 2249
 *Broderick, K., 2236
 Brodtkorb, F., 2269
 Broese, J. R., 2 81
 *Brugden, E. M., 2209
 2210
 Brugga, N., 2054
 *Brooks, R., 2240
 Brook, C., 2207
 *Brotherhood, P.,
 Ltd., 2251
 Brouha, A., 2182
 Brown, A., 2174
 *Brown, C. H., 2264
 Brown, D. J., 2022
 Brown, H. D., 2227
 Brown, J. C., 2043
 Bruce, W. F., 2023
 Bruckner, F., 2319
 *Bruder, K., 2105
 Bruckner, H., 2264
 Bruen, C., 2164
 Brummingham, L., 2278
 Brummann, B. W., 2031
 *Brulja, J., 2061
 *Brumbaugh, I. V., 2030
 Bruns, P., 2212
 Brush, W. W., 2216
 Brundlovskaya, A. E., 2197
 Bruun, J. H., 2275
 Bruns, P., 2114
 Bryan, A. B., 2092
 Bryk, 2021
 Bryllowa, L. P., 2197
 Bryland, G., 2140, 2195
 Buchholz, H., 2096
 Buchholz, H., 2121
 Buchholz, J. T., 2187
 Budewig, O., 2257
 Buchl, J., 2244
 Bucoer, H., 2015
 Bui, R., 2013
 *Bullington, R. M., 2218
 Bugakov, V., 2101
 *Bullant, T. A., 2231
 *Bull, H., 2111
 Buttol, J. K., 2180
 *Bumke, H. A. Ges., 2058
 Buerhard, B. F., 2043
 *Burd, J. L., 2312
 Burger, A. M., 2245
 *Burke, C. E., 2294
 *Burns, D. Jr., 2218
 Burlet, M. B., 2292
 *Burmeister, H., 2256
 Burett, F. M., 2184
 Burton, A. C., 2050
 Burton, D., 2224, 2225
 Busch, O., 2149
 Busch, M., 2121
 Bush, M. T., 2024
 Bushman, K. A., 2222
 Butcher, C. H., 2023
 Butler, A. O., 2066
 Butterworth, R., 2025
 Cable, D. R., 2385
 Caglioti, V., 2065
 Cain, R. S., 2121
 *Cain, J. R., 2105
 Cairns, R. W., 2034
 Cajola, G. G. R., 2244
 *Call, B. G., 2104
 Callison, W. B., 2179
 Calmon, H., 2072
 Calvery, H. O., 2125
 Cameron, D. H., 2021
 Cameron, R. P., 2235
 Campbell, W., 2094
 Campbell, W. B., 2246
 Cannato, L., 2180, 2197
 Cannon, J. G. Jr., 2214
 Cantoni, O., 2181
 Capper, N. S., 2172
 *Carlson Thoride Co.
 Ltd., 2251
 *Carcano, C., 2191
 Cardia, F., 2264
 Carey, B. W. Jr., 2189
 Carlisle, M. T., 2022
 Carlton, H. C., 2065
 *Carnahan, G. H., 2291
 *Caro, N., 2250
 Carpenter, L. V., 2319
 *Carr, C. B., 2109
 *Carr, R. H., 2275
 *Carreras, R. S., 2090
 Carter, R., 2257
 Carter, R. H., 2252
 *Carton & Papler-
 Ischik, G. Lager,
 2292
 *Casale, L., 2253
 *Casale Sacchi, M., 2024
 Cassano, A., 2194
 Casio, G. L., 2197
 Casella, U., 2182
 Casuburi, V., 2233
 *Cattelan, F., 2273
 Ceruti, G., 2191, 2195
 Chabrol, R., 2201, 2202
 Chadwick, J., 2048
 *Chagnaud, A., 2062
 Chakravorty, T. K., 2119
 Chah, P., 2093
 Challansonnet, J., 2095
 *Challenger, W. A. F., 2253
 Chalmers, S. W., 2118
 Chambers, W. H., 2190
 Chandler, W. F. Jr., 2056
 *Chapin, W. R., 2110
 *Chapman, W. B., 2274
 Charga, B., 2167
 Charidus, K., 2304
 Charonnat, R., 2201, 2202
 Chasnet, L., 2232
 Chaudron, G., 2092
 *Chemical Engineering
 & Wilton's Patent
 Furnace, Co. Ltd., 2275

- *Chemische Fabrik von Heyden A.-G., 2133, 2289, 2311
 *Chem. Fab. Kalk G m b H., 2251
 *Chem. lab. vorm. Sandoz, 2248, 2301
 Chertoukoff, R., 2041, 2132, 2159
 Childs, H. L., 2262
 *Chlowsky C., 2274
 *Chlison, H. I., 2306
 Chomkovic, G., 2174
 *Christensen, N. C., 2250
 *Christianson R., 2273
 *Chur, E., 2224
 *Chervikovskii, S. I., 2285
 Chwala, A., 2022
 Ciccio, I., 2199
 Cichocki J., 2042
 *"Culine-Wehr" Jewell Lorenz & Co., 2311
 Claus, R., 2216
 Claassen, H., 2232, 2320
 Claess, A. A., 2332
 *Clapp, A. L., 2289
 Clar, R., 2141
 Clark, G. L., 2049
 Clark, J. d. A., 2187
 Clark, K. A., 2278
 Clark, R. H., 2033
 *Clark, W. O., 2107
 Clarke, G. R., 2320
 Clarke H. T., 2117, 2154
 *Clarke, J. B., 2331
 Clausen, F. P., 2033, 2038
 *Clawson, M. S., 2103
 Clayton, E. S., 2204
 Clegg, E. L. G.
 Clement A., 2159
 *Clerc R., 2157
 Cleveland, R. E., 2055
 *Clifford A. M., 2332
 *Clifford W. B., 2030
 Cluzet, J., 2175
 Coblyn, J. H., 2031
 *Cochran J. M., 2318
 Cocks, H. C., 2056
 Codrime, P. W., 2286
 *Coc Mill Co., 2028
 Cohen, C., 2220
 Collie, W. M., 2117
 *Collie & Co., 2275
 *Collins, R. R., 2280
 *Collins, V. A., 2029
 Collins, W. F., 2103
 Colour Photographs (British & Foreign), Ltd., 2065
 *Colvin I. E., 2209
 Comet, M., 2181
 *Comolastic Corp., 2303
 *Compagnie de produits chimiques et électrométallurgiques Alais, Froges et Camargue, 2060
 Compere, E. L., 2196
 Comstock, G. F., 2047
 Conant, J. B., 2139
 Condores, P., 2199
 Conn, H. J., 2225
 *Consortium für elektrochemische Industrie G. m. b. H., 2251
 Constable, J. E. R., 2048
 *Contz & Huber, 2253
 *Coulage, W. D., 2027
 Coombe F. A., 2326
 Cooper, P. W. G., 2237
 *Cope, F. T., 2106
 Corbée, J., 2127
 Corlier, P., 2133
 Coria F., 2078, 2079
 Cornamusaz, P., 2140
 Corrigan J. F., 2245
 Corson, B. H., 2119, 2129
 Corson, M. G., 2099
 Cortese F., 2147
 Cosmulece I., 2183
 Costa A., 2195
 Coste, N. P., 2122
 Coster, G. De P., 2042
 Coughlin, E. D., 2248
 Coulson, H., 2329
 Coulter, T., 2261
 Cournot J., 2102
 *Courtauld, Ltd., 2273, 2248, 2290, 2305
 Coutance, G., 2122
 *Couture J. R., 2306
 *Cowdery, A. B., 2331
 *Cowell J. W., 2311
 *Cowper, E. O., 2331
 *Cowper-Coles, S. O., 2054
 *Coz, H. M., 2107
 Cox W. J., 2215
 Crabtree, J. I., 2045
 Crafalescu A., 2202
 Craig, P. H. C., 2285
 Craig, N., 2319
 Cramer, R., 2221
 *Crawford, F. A. E., 2253
 Crisera A., 2181
 *Cristalloff A.-G., 2246
 *Critchett J. H., 2108
 *Crocker, S., 2029
 *Crockett, J. B., 2332
 *Crosse, R., 2103
 *Crossley, P. B., 2259
 Crowther E. M.
 Cummings L. W. T., 2034
 Cunningham J. W., 2217
 Cuno, C. W., 2056
 Cuno I., 2049
 Curland E., 2269
 Currie W. A., 2221
 Curtis C. E., 2257
 Curtis W. F., 2051
 Custers J. P., 2050
 Cuthbertson, D. F., 2184
 Czerny, M., 2051
 Defert, G., 2171
 Dahle C. D., 2203
 Dalley, W. I., 2096
 *Dale J. G., 2027
 D'Amico A., 2206
 D'Amour, F. H., 2183
 Daniel, F., 2202
 Danile P., 2013
 Danner H., 2279
 Darbyshire O., 2031
 *Dargan W. H., 2105
 Darrow, D. C., 2174
 Darrow, M. A., 2229
 Darnin, C. G., 2048
 Dauteronde L., 2202
 *David, L. P., 2218
 Dvaidenkov, N., 2048, 2101
 Davidsohn, J., 2317
 Davidson, A. H., 2219
 Davies, D. T., 2182
 Davies, R. G., 2271
 Davies, W. L., 2315
 Davis C. P., 2031
 Davis, R. F., 2098
 *Davis J. C., 2106
 Davis, J. S. H., 2114
 Davis M. B., 2172
 Davis, P. A., 2379
 *Davis, S. H., 2030
 Day L. A., 2218
 De, S. C., 2118
 De Bartholomaeus E., 2271
 De Benedetti R., 2074
 Debré, R., 2185
 De Cacco E., 2240
 Decombe, J., 2117
 Deerr, N., 2319
 DeGuacomo U., 2199
 Dehns, R., 2193
 Dehn E., 2319
 Dekkers, W. A. M., 2050
 Dekker, M., 2116
 Delano A. J., 2197
 Delepaute, G., 2243
 Delfino, C., 2198
 De Marco, R., 2200
 Demmer A., 2101
 *Dengler A., 2306
 De Niederhäusern, 2193
 Denyer R., 2308
 *Dereks J., 2292
 De Riemas A., 2159
 Derrett Smith D. A., 2023
 Desmure G., 2327
 *Deslucias Drva D. D., 2155
 *Det norske Aktieselskab for Elektrokemisk Industri, 2060
 *Deutsche Bergbau A. G. für Kohle- und Erdölchemie, 2222
 *Deutsche Gold und Silber-Scheidanstalt vorm. Roessler 2030, 2060, 2109
 *Deutsche Hydrierwerke, A.-G., 2215
 Devoreux, W. C., 2078
 Devoto, G., 2042
 De Vries T., 2034
 Dezan, S., 2244
 Dhar, N. R., 2052, 2053
 *Diamond, C., 2303
 Dickie, H. A., 2096
 *Dickie, W. A., 2304
 Di Drilpas, S. D., 2230
 Dieke G. H., 2051
 *Dieterle, E. A., 2273
 Dietrich, W., 2128
 Dietzsch, O., 2161, 2162
 Dignaud, C., 2070
 Dilthey, W., 2065, 2127
 Dingemans, E., 2178
 Dinglage, R., 2127
 Dinslage, E., 2308
 *Dismiddle, J. G., 2301
 Dirks, B., 2229
 *Dirken, R., 2245
 Dlaney, G. S., 2308
 Dock, K. D., 2169
 Doan P. J., 2205
 Dörmberg, O., 2054
 Dohm, W., 2258
 *Dohrn, M., 2245
 Doktor, E., 2067
 Dolich, M., 2162
 Dolgov, B., 2124
 van Dommelen, A. M., 2047
 Dönnlag, K., 2320
 Donovan J. E., 2067
 Dony - Hensalt, P. (Hille), 2048
 *Dorner A. L., 2307
 Dorate R. W., 2032
 *Dorsey, F. M., 2110
 Dorta, C., 2110
 Dougherty, G. T., 2071
 Dowdell, R. L., 2100
 Downes, H. R., 2163
 Doyle, J., 2168
 Doson, T. P., 2220
 Drabkin, D. L., 2203
 Dragunov, S. S., 2247
 *Drake, J. L., 2258
 *Dreke, R. E., 2264
 Dreke T. G. H., 2174
 Dreger, E. E., 2119
 Drexel, E. G., 2197
 Dreyer, J., 2267
 Drexel, K., 2215
 *Dreyfus, C., 2306
 *Dreyfus H., 2152, 2155, 2156
 Drinker C. K., 2191
 Droné W., 2309
 Druce J. G. V., 2031
 Druzhinin S. I., 2260
 *Dubois R., 2061
 Dsh, E., 2093
 Düker, M., 2137
 Dürr, H., 2064
 *Dugas, L. J., 2256
 Dufogues, 2190

- Dumattresco - Monte 2190
 Dunaev, A. P., 2247
 Duninowski, A. I., 2073.
 *Dunker, H. C. L. 2332
 *Dunkley, H. M., 2251
 *Dunlop, C. K., 2304
 *Dunlop Rubber Co. Ltd., 2331
 Dunning, E. W. B., 2263.
 *Dunley, E. O., 2029
 *DuPont Visceloid Co., 2259
 Dupuy, E. L., 2046.
 Duren, 2221
 Durrant, T. H., 2041
 Durrant, F. J., 2029
 Dutcher, R. A., 2173
 Dyer, B., 2031
 Dyer, F. J., 2242
 Dyson, G. M., 2066.
 Eason, A. B., 2043.
 *Easton, O. J., 2061
 Eaton, P. M., 2233.
 Ebel, A., 2160
 Ebert, H. H., 2053.
 Ebert, L., 2043.
 *Eckell, J., 2060
 Eekl, K., 2230
 Eddington, A. S., 2046.
 Ederle, W., 2193
 Edington, O., 2223.
 Edwards, G. P., 2221
 Edwards, K. B., 2113.
 *Egan, J. J., 2109
 *Egler, N. P., 2105.
 *Eibes, B., 2234
 Eibner, A., 2309
 *Eichenberger, E., 2153
 *Eichwede, H., 2279
 "Eitracht" Braunkohlenwerke und Bricketfabriken, 2273
 *Eisenhut, O., 2060
 Eitel, H., 2193
 Ekkert, L., 2077
 Ekwall, A., 2262
 *Eldridge, E. B., 2056
 *Elfrith, H. D., 2222.
 Ehopoulos, S., 2202
 *Elia, C., 2312
 *Elia, C. H., 2303, 2303.
 Elia, E. T., 2257
 Elms, J. W., 2219
 *Elmen, G. W., 2111
 Elsd, J., 2043.
 Elphick, G. K., 2211
 Elson, L. A., 2146
 Elze, P., 2243.
 Emmermann, C., 2064, 2065.
 Emmons, R. C., 2031
 *E. M. S. Industrial Processes, Ltd. 2105
 Enders, W., 2083.
 *Engel, B., 2240
 Engel, L., 2164
 *Engel, O., 2291
 Engels, O., 2229
 Enna, P. G. A., 2327
 Epstein, A. K., 2316
 *Eskens, P., 2291
 Froukl, L., 2116
 Ferra, J., 2033
 *Erste Deutsche Kunstseilfabrik A. G. 2290
 *Escote, P., 2219
 *Eskesen, R. K., 2279
 *Esler, H., 2104
 *Etablissements Industriels de P. C. Grammont et de A. Grammont 2027
 *Etablissements Phillips & Pann 2223
 Evans, B. S., 2072
 Evans, O. J., 2121
 Evans, I. R., 2054
 Evans, W. L., 2114, 2115.
 Farnson, O. L., 2204
 *Fatt, W., 2236
 Eyles, A., 2102
 *Faber, H. B., 2315
 Faber, O., 2062
 Fahre, R., 2193.
 Fachgruppe für Brennstoff und Mineralien Chemie der Vereines Deutscher Chemiker, 2272
 Fachau, R., 2066.
 Faddersbill, H., 2177
 *Farber, E., 2290.
 Fahir, E., 2039
 *Farr, J. M., 2241
 *Falla, P., 2291
 *Fellon, J., 2106
 Fells, M. V., 2182
 *Felsale, C. m. b. H., 2236
 *Festing, J. S., 2263.
 *Favre, P. A., 2253.
 *Favre & Cie Zellenbetonfabrik Wallwien, 2263
 Fawns, H. T., 2163.
 Fedulov, N. S., 2073.
 *Feney, W. F., 2291
 *Fehrenbach, P., 2106
 *Feld, A. L., 2107
 Fenn, P., 2322.
 Fent, W., 2069
 Feldhaus, F. M., 2292
 *Feldhoff, E., 2029
 *Felix, F., 2299
 *Felsat, G., 2313.
 *Fenn, A. G., 2281
 *Ferd. Seibel Komm. Ges., 2302
 Ferguson, C. E., 2075.
 *Ferguson, J. E., 2332
 *Ferguson, J. & Sons, Ltd., 2332
 Fermann, W., 2023
 Fernor, L. L., 2082
 Ferrero, P., 2137
 Fichter, F., 2037
 Fieser, L. P., 2139
 Finckh, L., 2262
 Fink, H., 2026
 *Fink, M., 2030
 Finkelstein, H. M., 2165.
 Finn, C. P., 2271
 Finn, W. J., 2277
 Firket, V., 2084
 *Firth, L. T., 2103
 Fischlack, P., 2179
 Fischer, A., 2161
 *Fischer, A. C., 2263
 Fischer, I., 2178
 *Fischer, F., 2231, 2299
 Fischer, H., 2157
 *Fischer, H. G. M., 2274
 Fischer, M., 2219
 Fischer, O., 2163.
 Fischer, F., 2173
 Fischer, M., 2235
 Fleischer, F., 2192
 Fleuret, P. H., 2173.
 *Floris, H. G., 2103
 *Flowers, J. B., 2306
 *Folot, O., 2239
 Fong, H. C., 2213.
 *Follet, A., 2104, 2107
 Fomaz-Diaron, 2233
 Ford, S. G., 2119
 Forrest, H. O., 2031
 Forster, W. T., 2073
 Forstner, F., 2233
 Foshander, R. J., 2166
 Foster, L. S., 2170
 Foster, W., 2045
 Fournier, R., 2256
 Foa, C. J. J., 2244
 *Fox, J. H., 2253.
 Fox, P. S., 2221
 Fraas, F., 2217
 *Franchini, A., 2061.
 *Frank, A. R., 2230
 Franke, W., 2041.
 Franquon, 2270
 Freeth, F. A., 2041
 Freuch, H. J., 2104.
 *Frenkel, I. J., 2289
 Freyrejacquo, M., 2170
 Freiwirt, P., 2166
 Freudenberg, H., 2137
 Freudrich, H., 2307
 *Frey, C. N., 2239
 *Fried, K., 2023.
 *Friend, J. P. C., 2273
 Friend, W. H., 2207
 Fritz, H. R., 2329
 de Froberville, 2024
 Fröhlich, K. W., 2229
 Fuchs, W., 2123
 *Fues, E., 2291
 Fujm, Y., 2097
 Fukushima, M., 2065.
 Fulmer, H. L., 2078.
 Fulop, J., 2083
 Fulton, J. F., 2202
 *Funk, F. J., 2237
 *Furka, C. V., 2307
 Furutani, N., 2190.
 Gage, E. W., 2216
 Galecki, A., 2067, 2069.
 Galinsky, A., 2153.
 Galinovsky, P., 2149
 Galtner, K. W., 2172
 Gandrud, B. W., 2263.
 Gansel, G., 2181
 Gansser, A., 2322
 *Gant, H. P., 2223.
 Gardner, A. D., 2163.
 *Gardner, G. O., 2262
 Gardner, H. A., 2307
 Gardner, M. W., 2169
 Garner, T. L., 2329
 *Garrett, H. J., 2109.
 Gassner, G., 2234
 *Gathmann, E., 2106
 Gaudin, 2274
 Gaudin, A. M., 2052
 Gaudineau (Alim), 2231
 *Gautier, G., 2061
 Gavrilov, N. J., 2243.
 *Gebr. Palm Papierfabrik, 2292
 Geddes, R. C., 2270
 Gedroz, K. K., 2224
 *Gez, J. R., Akt.-Ges., 2301, 2329.
 *Gez, J. R. Soc. anon., 2300
 Geiger, H. L., 2044
 Gemant, A., 2047.
 *General Flec. Co., Ltd., 2254
 *General Zeolite Co., 2222
 Georgescu, I. D., 2174.
 Georgi, P., 2183
 *Gerlach, F. A., 2030
 Cerngross, O., 2324
 Gerry, L. S., 2292
 Gibson, C. S., 2113, 2117, 2121, 2146
 Carbenhain, H., 2037
 Gieser, H. S., 2032
 Gelbricht, R., 2072
 *Gillespie, H. R., 2313.
 Gelman, H., 2125, 2126, 2146
 Giordani, P., 2057.
 Giragosunts, G., 2197
 Glass, C. R., 2317.
 Glass, J., 2160
 Glasstone, S., 2056.
 Glud, G., 2269
 Glud, W., 2069, 2263
 Gnan, W., 2037.
 Goard, A. K., 2043
 Goche, O., 2043.
 Godbert, A. L., 2293.
 Görlacher, H., 2053.
 Götte, A., 2031
 Götte, K., 2297.
 *Goff, P. E., 2239
 Goldach, A., 2057
 Goldbeck, A. T., 2261
 Goldschmidt, P., 2269
 Goldstein, H., 2140
 Gooder, C. P., 2037.
 *Goodrich, R. J., 2154
 Goodway, N. F., 2140
 Gore, H. K., 2149

- Hopkinson, F. 2214.
Horn, L. F. 2214.
Hoeking, J. R. 2136.
Horn, W. 2092.
Hou, C. L. 2124.
Hosack, R. C. 2035.
Howard, N. J. 2219.
Howell, E. T. 2154.
Howitt, J. E. 2234.
Howlett, L. E. 2042.
Hoyt, S. L. 2110.
Hruska, J. H. 2092.
Hueh, C. M. 2114.
Hudson, R. F. 2094.
Hillemeyer C. 2023.
Hlad, L. C. 2279.
Hug, B. 2209.
Hug H. 2214.
Hult, P. 2274.
Hummel A. 2241.
Humphrey H. A. 2249, 2271.
Humphrey I. W. 2311.
Humphreys C. C. 2217.
Hunt, H. M. 2154.
Hunter, J. P. 2194.
Hunter, L. G. 2202.
Hunter, R. G. 2047.
Huththal, L. 2154.
Huthwecker W. 2127.
Hwyer, W. J. A. 2215.
Heda R. M. 2216.
I. G. Farbenindustrie, A.-G., 2023, 2059, 2090, 2109, 2152, 2153, 2154, 2155, 2156, 2157, 2215, 2236, 2246, 2249, 2250, 2251, 2252, 2253, 2254, 2256, 2263, 2264, 2273, 2274, 2280, 2281, 2284, 2289, 2290, 2291, 2292, 2294, 2299, 2300, 2301, 2302, 2303, 2304, 2305, 2306, 2307, 2311, 2313, 2325, 2331, 2332.
Imort, S. 2078.
Imperial Chemical Industries, Ltd., 2153, 2157, 2253, 2281, 2303, 2314, 2331.
Isaba, R. 2147.
Industriegas A.-G. Zweigniederlassung Wagram - Dissonagawerke, 2275.
Iones R. P. 2222, 2326.
International Industrial & Chemical Co., Ltd., 2251.
International Kautschukindustrie N. V. 2254.
International Fraepitation Co., 2062.
Ionesco-Craiova 2120.
Ionesco - Matia A. 2243, 2244.
Ionules, A. C. 2023.
Ipat'ev V. 2124.
Ipat'ev V. N. 2242.
Ippers J. W. 2064.
Irrera L. 2040.
Isakov F. N. 2230.
Ishino M. 2042.
Ismaïlov N. A. 2037.
Isobe H. 2090.
Isuki T. 2124.
Itiwara Y. 2045.
Itomi H. 2037.
Ivanoff D. 2138.
Ivanov P. P. 2141.
Ivanov N. 2160.
Iwamoto K. 2132.
Iwas K. 2045.
Iwaya, K. 2147.
Jacob K. D. 2229.
Jacobs, A. A.-G., 2319.
Jacobs P. 2330.
Jacobson K. P. 2142.
Jacobus D. S. 2027.
Jacobs G. T. 2029.
Jadhav G. V. 2125.
Jadoul A. 2274.
Jaeck, W. 2279.
Jachowky J. J. 2279.
Jakovets, P. 2232.
Jambhewala G. B., 2134.
James, C. 2245.
James R. G. 2331.
James, T. G. I. 2142.
v. Jacob M. 2154.
Jarrier, P. 2024.
Jarvis, B. W., 2174.
Javiliet, M. 2208.
Jebson - Marwedel, H. 2256.
Jelicks, O. 2221.
Jelinsk, H. 2035.
Jenkins, C. H. M., 2091.
Jenne L. L. 2221.
Jennings G. W. 2315.
Jennings J. M. 2241.
Jewel R. 2275.
Jewell, W. R. 2233.
Jimenez, M. E. 2164.
Jochmann, E. 2164.
Jone, C., 2260.
Johannessen, F. 2203.
John, P. 2141.
Johnson, C. R. 2070.
Johnson, J. D. A., 2113.
Johnson, L. 2111.
Johnston, F. R. 2233.
Johnston, H. W. 2234.
Johnston, M. W., 2183.
Jollot P. 2042.
Jolson L. 2207.
Jonas, K. G. 2284.
Jones L. C. 2249.
Jones L. H. 2025.
Jones, R. L., 2168.
Jones, R. S. 2223.
Jones T. V. 2043.
de Jongh, E. E. 2174.
Jordan H. 2157.
Jouling C. H. 2070.
Jorns G. 2175.
Josephson K. 2142.
Jouan 4 2200.
Journiaux A. 2074.
Joussan P. 2242.
Jungwirth O. 2099.
Junkerdoerf P. 2174.
Jost 4 2027.
Kadner P. 2034.
Kadow A. 2254.
Katzel H. 2124.
Kahbaum W. 2109.
Kahlenberg L. 2117.
Kalinowski J. 2191.
Kaiser Wilhelm Institut für Eisenforschung V. 2074.
K&S Chemie A.-G. 2250, 2251, 2252.
Kallam P. L. 2277.
Kallander O. 2257.
K&L & Co. A.-G. 2074, 2249.
Katschew D. 2277.
Kameda T. 2074.
Kamerman, P. 2226.
Kamohata Y. 2225.
Kane H. 2104.
Kansob, A. 2311.
Kao C. H., 2134, 2171.
Karnahl H. 2274.
Karrer P. 2203.
Karström, H. 2181.
Karyanova N. 2317.
Kastner, R. 2291.
Kato, T. 2132.
Katoh, N. 2099.
Kaufer, P. 2135.
Kauro, B. D. 2074.
Kavtzevich N. N., 2023.
Kawai, T. 2044.
Kawata S. 2042.
Kayser H. 2054.
Kazantseva V. A. 2270.
Kran R. H. 2217.
Kray, H. O. 2244.
Kreman, J. H., 2034.
Kremer P. 2041.
Kehr R. W. 2219.
v. Keil, O. 2099.
Kell W. 2175.
Keltel, W. 2024.
Keth, J. E. 2205.
Keller, K. 2269.
Kelley, F. C. 2110.
Kellner L. 2173.
Kellner A. P. 2093.
Kemp A. R. 2230.
Kempter, K. 2032.
Kenda H. 2301.
Kennedy, C. C. 2301.
Kennedy, J. C. 2274.
Kenyon J. 2130.
Kern, H. M., 2191.
von Kerpely K., 2054.
Kerr P. R., 2264.
Kershaw J. B. C. 2267.
Kellins A. 2073.
Kethel R. J. B. 2079.
Keyes F. G. 2034, 2037.
Kell P. 2209.
Klech V. C. 2179.
Klemm W. 2170.
Kilp W. 2237.
Kimla H. S. 2247.
Kindwer P. 2304.
King J. G. 2253.
King J. G. 2264.
Kinney S. P. 2044.
Kirchhof A. 2104.
Kirchhuth 2025.
Kistara S. 2126.
Kirchleiss, I. 2255.
Kim A. 2103.
Kistner A. J. 2244.
Kitchin D. W., 2277.
Klaß L. 2024.
Klapheck, J. 2047.
Klarer J. 2153.
Kleemas, 2227.
Kleman, R. D., 2051.
Klema, P., 2252.
Klumpt, W. 2262.
Klma, J. 2031.
Klog, M. 2224.
Klönner H. W. 2026.
Klönner P. W., 2033.
Kloppstock P. 2182.
Kluger W. 2030.
Klui, H., 2227.
Kluniewicz, P., 2202.
Klaus P. V. 2059.
Knerr, H. C., 2090.
Koblinger, G., 2304.
Koblinger, L., 2304.
Knoll W. V. 2291.
Knolle W. 2027.
v. Knorr K. 2167.
Kob, H. 2251.
Koch E. 2321.
Koepfel P. 2219.
König, J., 2208.
Königs W., 2294.
Köster, W., 2094.
Kolman, T. 2175.
Koran, A. I., 2304.
Kobman, E. P., 2173.
Kozumi, M., 2027.
Kolhaiker, G. B. 2126.
Kolbbrunner, H. 2264.
Koller, L. R., 2054.
Kolthoff, I. M., 2042, 2076.
Komatsu T., 2273.
Kon, G. A. R., 2137.
Kondo, H., 2147.
Konen, H., 2054.
Kosajmans, J., 2172.
Kopelina, A. W., 2243.
Kopp P., 2241.
v. Korányi, A., 2194.
Kordas, E., 2043.

- *Korbeck, G., 2273
Kosakovich, P. F., 2007
Kosko, Y. D., 2190
Kosvsky, M., 2273
Koyanagi, K., 2260
Krack, F. C., 2037
Kraeber, L., 2082
Kramer, K., 2179
Kramer, P. H., 2110
Kramers, H. A., 2023
Krauss, F., 2260
Kraybill, H. R., 2169
Kraha, O., 2115
Kraus, H., 2261
Krack, A. E., 2030
Kriehberg, T., 2027
Krich, H., 2193
Kriemsson, A., 2022
Kron, L. C., 2035
de L. Kroon, R., 2046
Kronstad, H., 2263
Krod, H., 2154
Kruithof, A. A., 2050
Krupp, F. A.-G., 2109, 2107, 2105
Krum, 2219
Krustovs, J., 2033
Kubacka, V., 2222
Kucharska, P., 2223
Kuka, A., 2164
Kuka, R., 2164
Kuntzel, A., 2161, 2162, 2227
Kupper, A., 2192
Kupelmann, L. N., 2155
Kupelmann, B., 2193
Kuba, F., 2241
Kuba, F., 2245
Kuka, R., 2153
Kunzel, H. P., 2220
Kukla, O., 2253
Kunstlinger - Pat. ent. - Verwertungs A.-G., 2236, 2242
Kupka, J., 2219
Kupka, K., 2034
Kurokawa, T., 2190
Kurra, M., 2102
Kusnetsov, E. A., 2079, 2080
Kyrassov, N. C., 2053
Labe, A., 2244
Lacey, R., 2113
Lacour, A., 2090
La Fave, P., 2240
Lafente, F., 2221
Lafitte, P., 2293
Lagata, H., 2234
Lahay, F. T., 2206
Lahle, T., 2106
Laud, W. G., 2240
Lake, C. H., 2045
Lamb, M. C., 2205
Lama, 2085
Lamont, N. E., 2219
Lampe, E., 2227
Landau, R., 2053, 2060
Landabere, F. E., 2200
Landes, C. H., 2253
Landes, W., 2279
Lang, H. R., 2275
Lang, E., 2041, 2057
Langmeyer, C., 2239
Langruth, S., 2112
Langner, A., 2096
Lania, M., 2235
Larson, L., 2106
Lashmann, H., 2031
Laudet, G., 2045
Laurson, P. G., 2215
Laut, K., 2090
Laux, J., 2210
Lawrence, M. L., 2255
Lazarov, N. V., 2197
Lazarov, P., 2250
Lazarovich, N. A., 2073
Leathers, J. B., 2157
Leaver, E. S., 2052
Le Bonet, L., 2061
Lecomte de Noüy, P., 2190
Lecon, R., 2204
van Leeuwen, M., 2171
Lelebert, A., 2264
Lefèvre, J., 2274
Lepier, E., 2251
Leyer, E., 2141
Ley, B. B., 2253
Ley, V. E., 2106
Leyre, M., 2264
Lehmman, P., 2204
Lehrke, H., 2236
Lugnos, R. B., 2215
Linn, H. S., 2254
Lerner, H., 2245
Leyens, L. S. M., 2251
Lendtayer, E., 2149
Lemane, R. M., 2027
Lemel, F. B., 2040
Lenn, L., 2140
Lepner, W., 2208
Lepruvak, W., 2063
Levranc, A., 2044
Levi, A., 2194
Levi, C., 2277
Levi, M. G., 2264, 2277, 2271
Levin, H. L., 2264
Lewcock, W., 2137
Levin, A. T., 2223
Lewis, I. M., 2167
Leibovitch, G. L., 2255
Leichtenstern, R., 2173
Liddle, E. C. T., 2227
Leibrecht, E., 2092
Lehr, 2223
van Leempt, J. A. B., 2091
Lengyang, H., 2229
Lengyang, R. E., 2042
Leimbach, N., 2151
Lensch, O. N., 2223
Lend, J. O., 2190
Lendhard, F. T., 2262
Leng, A. W., 2223
Lerner, H., 2042
Lentzen, C. F., 2070
Linton, R. R., 2183
Llowsanika, S., 2256
v. Lippmann, E. O., 2121, 2119
Lison, L., 2182
Liss, G., 2192
Liss, H., 2029
Litavsky, G., 2183
Littlejohn, W., 2278
Litterbach, P. M., 2207
Livingstone, A. V., 2122
Ljushberg, K., 2091
Lloyd, D. J., 2122
Lloyd, J. W., 2232
Lockmann, G., 2175
Locke-Holmgren, W. L., 2031
Lodge-Centrel, Ltd., 2061, 2062
Lodge, L., 2061
Lodge, O., 2054
Lof, K., 2215
Loebner, A., 2173
von Loewke, H., 2207
Loewer, A., 2193
Loewenthal, J. F., 2229
Loew, G., 2215
Löwenstein, E., 2070
Loewenstein, N. L., 2277
Loewenstein, Z. V., 2222
Lohmann, K., 2173
Lomas, R., 2251
Loewenstein, G. G., 2073
Lorenz, W. F., 2057
Loren, A., 2273
Loren, M., 2273
Loren, W., 2273
Loren, H., 2197
Loverin, L. D., 2030
Lower, M. S., 2231
Lowry, T. M., 2097, 2147
Lowry, H., 2132
Lozier, J., 2162
Labman, N., 2029
Labowsky, S. J., 2222
Laba, H. A., 2001
Lucas, F. V., 2099
Lucas, O. D., 2279
Lucas, R., 2072
Lucas, W., 2201
Luchman, R., 2176
Lusk, J. M., 2179
Lackebach, R., 2104
Lackow, J. H., 2079
Ludvik, F., 2291
Lieber, H., 2209
Lieber, O., 2115
Lieber, E., 2227
Lippe-Cramer, 2064
Layon, J., 2234
Lukas, A., 2182
Lundborg, M., 2204
Lundberg, A., 2262
Lury, M., 2214
Luther, M., 2223
Luyken, W., 2092
Lux, H., 2111
Lycas, W. H., 2114
Lyon, A. V., 2233
Lyubimova, M. P., 2197
Maas, E., 2092
Maas, O., 2123
Ma, S.-Y., 2134, 2135
McAlister, A. S., 2214
McArthur, C. B., 2155
McBride, B. V., 2059
McCaas, S. E. O., 2169
McCarthy, J. C., 2213
McCanley, H. J., 2090
McClure, E. F., 2106
McDonald, D., 2022
McDonnell, R. E., 2222
McFarlane, W. D., 2078
McGeorge, W. T., 2226
McGovera, W. J., 2254
Mack, F., 2233
McHargue, J. S., 2207
McLaws, P., 2212
MacKenzie, L. R., 2242
Mackie, T. J., 2154
McKinley, C. W., 2025
MacKinnon, D. A., 2030
McLaughlin, A. R., 2179
McLennan, J. C., 2047, 2050, 2031, 2095, 2099
McLeod, J. H., 2047
McLaughlin, T. J., 2219
McMichael, P., 2267
MacNaughton, A. L., 2237
McPherson, W., 2031
McQuaid, H. M., 2090
Machmow, H., 2159
Marbach, M., 2219
Mayer, E. H. J., 2031
Magnet Work Ges. Eisenach Special Fabrik für Elektromagnetapparate, 2259
Magnus, A., 2037
Mahood, S. A., 2125
Manner, F., 2176
Marano, M., 2157
Marzels, M., 2185
Makrolin, R. L., 2003
Maltherte, I. de V., 2225
Mabutra, R. C., 2227
Malinovich, V. E., 2076
Malberry, J., 2251
Malych, G., 2106
Mamchenkov, I. P., 2231
Manabe, T., 2195

- *Manchester Oase Co., Ltd., 2304.
 Mandl, K., 2261.
 de Mandrot, G., 2041, 2139.
 *Manes, A., 2242.
 Manfred, O., 2210.
 *Manfred Weiss Stahl- und Metallewerke, A.-G., 2239.
 Manpiu, A., 2316.
 *Manson, G. N., 2209.
 Meun R. F., 2034.
 *Meun, W., 2232.
 *Mante F. F., 2106.
 Meun L., 2151.
 Marcusson, J., 2270.
 Marlon, P., 2157, 2243.
 Margraf, I., 2120.
 Mare, C., 2037, 2040.
 Marresco, N., 2039, 2040.
 Markham E. C., 2039.
 Martoth, R. H., 2170.
 *Marquardt, A., 2206.
 Marrack, J. R., 2143.
 Marrasini, A., 2144.
 Marvotti, R. H., 2223, 2224.
 Marshall, C. E., 2039.
 Marshall, F. W., 2142.
 Martell, P., 2314.
 Martin, F. T., 2041.
 Martin, G., 2223.
 *Martin, W. E., 2106.
 Martus, E., 2199.
 Marvel, C. S., 2113, 2114, 2116.
 *Marx, K., 2304.
 Mascherpa, P., 2194.
 *Maschinenfabrik Augsburg Nürnberg, A.-G., 2292.
 *Maschinenfabrik vorm. G. Dorst, A.-G., 2159.
 Mascru, M., 2163.
 Maskell, E. G., 2169.
 Mason, F. A., 2158.
 Mason, P. G., 2169.
 Mason, H. L., 2166.
 Mason, R. L., 2144.
 Mason, T. N., 2293.
 Massey, H. S. W., 2049.
 Matsumo, D., 2316.
 Matukas, P., 2176.
 *Mathews, H., 2108.
 *Mathews, W., 2108.
 Mathews, J. A., 2102.
 Metchky, 2220.
 Matsui, M., 2073.
 Matsumoto, H., 2196.
 Matthe, H., 2173.
 Meume, L., 2234.
 *Maurel Investment Corp., 2272.
 Mevrodin, A., 2116.
 Maxwell, L. C., 2203.
 Maxwell, L. R., 2044.
 Maxima, M., 2201, 2202.
 Mayer, K., 2024.
 *Mayer, M., 2273.
 Maynard, C. E., 2339.
 *Mechanite Metal Corp., 2107.
 Meness, H. S. V., 2014.
 Menicke R., 2164.
 *Meis, H., 2134.
 Meisse, L. A., 2102.
 Meissner W., 2033.
 *Meibhardt, H., 2024.
 *Meish, T. G., 2029.
 Meisner, A., 2034.
 Melzak, J., 2190.
 Mendel, B., 2176.
 Meneghetti, R., 2166.
 *Menk, W., 2192.
 Menton J., 2204.
 Merck E., 2014.
 Mereskov, B. V., 2079.
 Merkel, F., 2023.
 *Merkel & Klein Co., 2303.
 Merri L. P. H., 2223.
 Merri H. B., 2223.
 Merz, K. W., 2132.
 Meissner M., 2199.
 *Metallgesellschaft, A.-G., 2061, 2062, 2104, 2219, 2272.
 van der Neulen, P. A., 2034.
 *Meyer, C. H., 2209.
 Meyer, D., 2230.
 Meyer - Dietz, G., 2140.
 *Meyer, J. W., 2254.
 Meyer, K., 2047.
 Meyer, P., 2133.
 Meyerhof, O., 2177.
 *Meyerhofer, A. P., 2101.
 *Meyers, S. E., 2029.
 *Meyersberg, P., 2233.
 Michelsen, J., 2178.
 Michot, P., 2040.
 *Midden Europäische Oefroomaatschappij, 2231.
 *Mieder, F., 2025.
 *Mietzsch, F., 2133.
 Mihul, C., 2044.
 Milas, N. A., 2119.
 *Miller, H. C., 2233.
 Miller, H. K., 2203.
 *Miller, T., 2306.
 Millwood, J. P., 2214.
 Milne, D., 2171.
 Milne, V. E., 2204.
 Min, P., 2193.
 *Minkoff, G., 2290.
 Minomura, T., 2163.
 Munster, C. C., 2277.
 Muschnet, H., 2176.
 Mushchenko, K. P., 2040.
 Mitchell, C. A., 2237.
 Muelo, M., 2200.
 *Mitscherling, W. O., 2291.
 Miyaji, K., 2238.
 Miyamoto, S., 2042.
 Mochl, O., 2270.
 *Mocker, C., 2303.
 Moehlig, R. C., 2147.
 Moeller, 2230.
 Moer, H., 2043.
 *Moeller, M. E., 2024.
 *Moeller-Moeller, 2219.
 Moers, L., 2234.
 Moers, K., 2038, 2043.
 Moffitt, W. G., 2131.
 Mogensen E., 2177.
 Mohrman F. W., 2221.
 Moir, G. M., 2203.
 *Mooforts, F. A., 2308.
 Monheim, J., 2041, 2037.
 du Mont H., 2114.
 *Moontershof, societ  generale per l'industria mineraria ed agricola, 2230.
 Moontgomery, B. E., 2167.
 Montgomery R. J., 2237.
 Montague E., 2131.
 Monton, V., 2101.
 Moore, D. K., 2144.
 Moore E. J., 2234.
 *Moore, R. L., 2331.
 *Moore, R. P., 2029.
 Moravci, E., 2200.
 Morales - Otero, P., 2168.
 Moriel, H., 2216.
 Morosiewicz, J., 2060.
 Morris, H. P., 2071.
 *Morrow, J. H., 2213.
 Morse, J. K., 2162.
 *Mortensen, C., 2209.
 Moser, M., 2087.
 Moskowitz, S., 2067.
 Moss, J. A., 2073.
 Motchman, F. W., 2254.
 Motter, M., 2137.
 Moutseron, 2078.
 Mowry, H., 2209.
 Moyers, 2215.
 Moyer, W. W., 2113.
 Moys, H., 2063.
 *Moyer, A., 2233.
 Muehlberger, G. W., 2201.
 M    , 2264.
 M    , A., 2119.
 *M    , I., 2027.
 M    , P., 2164.
 M    , R., 2204.
 *M    , W. J., 2232.
 Munster, F., 2228.
 M    , W., 2224.
 M    , G., 2104.
 *Munster, D. A., 2027.
 M    , R. S., 2031.
 M    , M., 2196.
 *M    , E. O., 2306.
 Muntwyler, E., 2191.
 Murakami, T., 2097, 2098.
 Muraour, H., 2293.
 *Murphy, A. R., 2133.
 *Murphy, E. A., 2331.
 *Murray, H. D., 2063.
 Murray, R. K. S., 2234.
 Maus, J., 2042.
 *Myers, E. W., 2104.
 *Myers, J. W., 2030.
 Myers, V. C., 2191.
 *Mylo, O., 2060.
 *N. V. de Bataafsche Petroleum Maatschappij, 2156, 2272.
 *N. V. J. A. Carpy's Garenfabrieken, 2249.
 *N. V. Maatschappij tot exploitatie van Veredelingsprocedures, 2234.
 *N. B. Nederlandse Keuntdielefabriek, 2249.
 *N. V. Philips' Gloeilampenfabrieken, 2109.
 *N. C. W. Paot & Varush Remover Co., Ltd., 2234.
 Negel, R. H., 2204.
 Negel, W., 2054.
 Nahum, L. H., 2190.
 Nakamura, S., 2140.
 Naom, F., 2233.
 Nash, H. E., 2294.
 *National Carbon Co., 2022.
 *National Paper Process Co., 2292.
 *Natta, G., 2273.
 Naom, F., 2162.
 *Neuk, J. J., 2026, 2027, 2221.
 Neumann, K., 2029.
 *Neal, J. R., 2210.
 Nekryt, S. S., 2098.
 Neftestan, F. J., 2278.
 Neles, L. H., 2319.
 Nelson, J. W., 2071.
 Nestoresco, B., 2153.
 Neugebauer, 2064.
 *Neuls, J. D., 2236.
 Neumann, G., 2270.
 Neumeier, F., 2132.
 Neuburg, L. H., 2153.
 New, G. F., 2303.
 Newcomb, C., 2164.
 Newell, W., 2309.
 Newens, F. R., 2065.
 Newman, L. F., 2320.
 Newton, R. C., 2314.
 Newton, R. F., 2034.
 Newton, W., 2233.
 Nichols, S. P., 2169.
 Nieder, J. B., 2077.
 Niemer, H., 2179.
 Nieschitz, O., 2192.
 Nigh, P., 2262.
 Niklas, H., 2278.
 Nikolaev, O. V., 2178.
 *Nikolai, F., 2239.
 Nishida, K., 2242.
 Nishigori, S., 2100.

- Rebek, M., 2045
 Redl, R., 2187
 Reed, M. C., 2331
 Rees, W. J., 2253
 Reetz, A., 2263
 Reichardt, H., 2050
 Reid, A., 2043
 Reif, C., 2077
 Reihing, H., 2234
 Reilly, A., 2290
 Reimer, W., 2025
 Reinhardt, F., 2257
 Reiss, K., 2030
 Reiter, E., 2083
 Remy, J. T., 2216
 Remly, F. O. F., 2149
 Rentachler, H. C., 2027
 Repovš, B., 2091
 Reschke, J., 2174
 Reuss, C., 2315
 Rewald, B., 2207
 *Rewald, B., 2127
 Reyschler, A., 2063
 Reynold, P. B., 2224
 *Rhodes, P. O., 2273
 Rhodes, P. H., 2023
 Riccardi, L., 2230
 *Ries, W. S., 2061
 *Richards, W. A., 2110
 *Richardson, L. T., 2072
 *Richardson, W. W., 2062
 Richter, C., 2223
 *Richter, G. A., 2291
 *Richter, K., 2110
 Riede, W., 2207
 *Riedhammer, L., 2029
 *Rieger, E., 2155
 Riez, H., 2304
 Riese, W., 2090, 2269
 Riezler, W., 2046
 Riffart, H., 2164
 Rindolf, L., 2295, 2296
 Rinse, J., 2309
 Risch, D. McK., 2202
 Rios, J. R., 2200
 *Ripke, O., 2235
 Rischbuth, P., 2045
 *Risler, J., 2061
 Riwel, R., 2260
 *Ritter, J. G., 2111
 Rivoob, P. I., 2163
 Robert, J., 2295, 2296
 *Roberts, P. G. L., 2106
 Roberts, W. M., 2201
 Robertson, J. M., 2096
 Robertson, M. E., 2322
 Robn, G., 2063
 Robinson, R., 2149
 Roche, J., 2177
 Rockwell, G. E., 2322
 *Roddy, G. R., 2223
 *Rodgers, R. L., 2279
 *Rodolfo, E., 2251
 *Roeder, V., 2027
 Röhring, A., 2234
 *Röhn, O., 2232
 Röhrig, H., 2093, 2103
 Rölt, E., 2119
 *Rogers, P. P., 2118
 Rogers, J. S., 2324
 *Rogers, T. H., 2279
 Roginski, S., 2053
 *Rogler, A., 2310
 Roll, P., 2034
 Roll, L. J., 2110
 Romashkevich, I. P., 2231
 Rona, P., 2160
 Rouchesse, P., 2080
 *Rondek, G., 2292
 Roodenburg, N. M., 2274
 Rone, J. O., 2261
 Roseberry, H. H., 2162
 Rosen, B., 2109
 Rosenbaum, F., 2204
 Rosenhauer, 2094
 Rosenkovich, I., 2053
 Rona, F. P., 2063
 Rona, M. H., 2054
 Rona-Rona, D., 2247
 Rossmann, F., 2309
 *Roth, C., 2060
 Roth, W. A., 2093
 Rothe, J., 2277
 Rothe, O., 2073, 2103
 *Rothers, W. S., 2111
 Rothmann, H., 2191
 Rotgardt, A., 2167
 Rottzehl, E., 2123
 Rottzehl, L., 2234
 Rowe, L. P., 2023
 Rowinski, P., 2141
 *Rowton, H. A., 2236
 Roy, W. R., 2207
 Royen, P., 2043
 Rozentzels, E., 2041
 Rubens, S. D., 2276
 Ruby, A., 2244
 *Rudin, E., 2244
 Radolph, H., 2047
 *Ruf, A., 2316
 Ruff, G., 2041
 Ruggli, P., 2294
 *Rule, J. P., 2235
 *Runway, H. S., 2027
 Rupp, H., 2046
 Rupp, H., 2174
 *Russ, E. F., 2060
 Russinova, K. I., 2162
 Russo-Bonds, G., 2199
 Russo, G., 2181
 Runzick, L., 2156
 Sabetta, V. J., 2075
 Saccardi, P., 2200
 Sachs, G., 2086
 de Sacy, G. S., 2182
 *Saecker, C. M. Jr., 2061
 Sab, P. P. T., 2135
 *Sailer, W., 2163, 2328
 Salhard, E., 2319
 *Salanderichin, N., 2104, 2107
 St. John, N. B., 2146
 Sakata, S., 2195
 Saladin, B., 2276
 *Salenius, P. G. N., 2209
 *Salenius, E. S., 2209
 Salomone, G., 2297, 2316
 Salomonsen, L., 2196
 Samethima, J., 2079
 Sampletin, G., 2172
 Samuel, F. L., 2143
 Sanchez, J. A., 2213
 Sander, A., 2308
 *Sander, F., 2023
 Sanders, K., 2264
 Sanders, G. P., 2293
 Sando, C. F., 2131
 Sankaran, O., 2164
 Sankar, C. A., 2123
 Sanson, I., 2240
 Sapozov, E., 2194
 Sarre, H., 2179
 Sartori, A., 2077, 2194
 Savaki, Y., 2077
 Satoh, S., 2092
 Sauer, J., 2144
 Savage, C. G., 2231
 Saville, C. M., 2217
 Saxon, R., 2042
 Savitsky, V., 2310
 Scales, O. L., 2266
 Scarpa, O., 2057
 Scarveth, G. D., 2223
 Satchard, G., 2034
 Schaffner, P. V. I., 2123
 Scharian, H., 2172
 von Scheele, H. K., 2093
 Scheffer, P., 2229
 Scheffers, H. W., 2315
 Scheffau, L., 2075
 Scheide, G., 2070
 Scheiber, H., 2307
 Scheiff, W., 2178
 Schelmski, N., 2179
 Schelmski, N., A., 2070
 Schenk, P. W., 2026
 *Schering, K. H., 2026
 A.-G., 2157
 Scherr, K., 2189
 Scherrert, A., 2174
 *Scheyer, H., 2300
 *Schiffert, C. J., 2029
 Schilling, C., 2194
 *Schlappert, H., A., 2060
 *Schleich, H., 2026
 Schlumpf, E., 2245
 Schmidt, E., 2035
 Schmidt, A., 2075
 Schmidt, A. A., 2161
 Schmidt, C. L. A., 2042
 Schmidt, H., 2243, 2252
 Schmidt, K., 2131
 Schmidt, M., 2057
 *Schmidt, W., 2109
 *Schmidt'sche Heissdampf G. m. b. H., 2025
 *Schmidt, E. M., 2062
 *Schmeyer, E., 2030
 Schmitt, F. O., 2025
 Schmitt, O. H. A., 2025
 Schmitt, R., 2085
 Schmitt, P. M. E., 2276
 *Schneebell, J., 2223
 *Schneible, C. B., 2029
 Schuetter, O., 2070
 Schnoutka, 2066
 *Schuch, E. P., 2232
 *Schueller, W., 2155, 2157
 Schoenen, H., 2175
 Schoonmaker, P., 2058
 *Scholz, V., 2313
 Schommer, W., 2127
 Schopper, L., Labors stories, 2243
 *Schradler, H., 2153
 Schrauth, 2112
 Schroeder, R. P., 2114, 2118
 Schrüfer, J., 2269
 Schubert, P., 2025
 Schuch, W., 2149
 Schuler, H., 2030
 *Schuettgenheim, J. B., 2105
 Schults, A., 2164
 Schults, L. H., 2041
 Schultze, K. W., 2186
 *Schultze, W., 2312
 Schults, M. H., 2092
 Schultze, A., 2031
 Schultze, P., 2146
 Schultze, K., 2294
 Schumb, W. C., 2071
 Schwartz, A. M., 2024
 *Schwartz, H. A., 2106
 *Schwartz, K. W., 2050
 v. Schwarz, M., 2091
 Schwarz, R., 2047, 2309
 *Schweitzer, H., 2299
 Schwilung, W., 2099
 Schwob, M., 2052
 Scott, H., 2049
 Scott, R. B., 2021
 *Scott, R. H., 2110
 Scott, H., 2150
 *Scout Co., The, 2292
 Sears, P. E., 2046
 Sebesta, W., 2048
 *Seibell, L. D., 2132
 Seekamp, H., 2090
 Seemann, H. J., 2033
 Segal, B., 2205
 Seifan, M., 2024
 Sekiguchi, H., 2097
 *Selenophon" Licht- und Tonbildgewerkschaft m. b. H., 2038
 Seligberger, L., 2324
 Serantes, M. T., 2077
 Serbriakov, A. M., 2297
 Sergeev, A., 2073
 *Sergeev, R., 2104
 Severino, A., 2199
 Seri, T., 2048
 Seyer, W. P., 2040
 Seyewetz, A., 2062
 *Slovak" Společnost Smaltony a Továrny na Kovové Zboží Akc. Spol., 2029

- *Shackleton, W., 2030
Shandorov, A. M., 2073
Sharp, D. F., 2256
Sharp, G. S., 2163
*Sharpless Specialty Co., 2281
Shaughnessy, J., 2191
*Shaw D. N., 2332
Shaw, H., 2140
Shchepov N. P., 2087
Schcherbakov, A. P., 2232
Shread, A. C., 2071
Shemyakin, F. M., 2032
Sheppard S. F., 2044
*Sheppard S. E., 2065
Sherman H. C., 2158
Sherrill M. S., 2030
*Shinn E. W., 2110
Shishido H., 2147
Shorey, P. C., 2224
Shriner, R. L., 2119
Shukla, S. N., 2077
Siebert, I., 2142
Siebert Modrow, I., 2189
*Siekani, G., 2191
Siegel R., 2193
*Siemens - Bauwesen G. m. b. H. Kotum.-Ges., 2060
*Siemens & Halske A.-G., 2027, 2060, 2078
*Siemens - Schuckert werke A.-G., 2060, 2061, 2274
Sierakowski, S., 2190
*Signum A.-G., 2061
Silberkweit E., 2151
*Silver Springs Bleach log & Dyeing Co., Ltd., 2302
Simeci D., 2177, 2202
Simon, A. W., 2058
*Simon M., 2290
*Simon, M. E., 2062
Simonin G., 2248
Simonsen, J. L., 2121, 2134
Simpson, B. W., 2173
Simpson S. G., 2071
Simon, F. W., 2201
Singer, P., 2257
Singer L., 2276
Smith B., 2128
Smitamed C., 2264
*Smyth, P. S., 2253
Skau E. L., 2023
Skinner, C. E., 2228
*Skoda Works, Pizen 2106
Slabber M. H., 2226
Slotte K. H., 2042, 2133
*Smallwood A., 2106
Smekal A., 2051
Smith A. P., 2179
Smith A. S., 2329
*Smith C. M., 2102
Smith, E. L., 2034
Smith, G. F., 2071
Smith, H. D., 2051
*Smith, H. F., 2274
Smith J. C., 2127
Smith J. H. C., 2112
Smith, L. B., 2038, 2037
Smith L. I., 2128, 2125
*Smith, M. J., 2304
Smith, R. G., 2203
*Smith W. G., 2093
*Smith, W. H., 2110
*Smith W. N., 2059
*Smith, W. S., 2104
*Smithson, J. K., 2085
Smits A., 2045
Smyth C. P., 2032
Smythe C. V., 2045
*Soell, F. D., 2264
Sobruho, A. P., 2073
*Soc. anon. des accumulateurs monoplaque 2058
*Soc. anon. pour l'ind. chim. à Bâle 2028, 2060, 2245, 2249, 2253, 2288, 2302, 2310
*Soc. anon. Le Carbone, 2055
*Soc. anon. d'Ouvert.-Misebaye, 2028
*Soc. anon. des verres & bouteilles du Nord, 2258
*Soc. d'applications et de recherches scienc. liquides et industrielles (S. A. R. L.), 2306
*Soc. chimique de la grande - paroiase Azote et produits chimiques 2236
*Soc. d'électrochimie, d'électro-metallurgie et des acières électriques d'Ugine, 2104
*Soc. française de catalyse générale, 2135
*Soc. de Prospection Electrique (Procedés Schlumberger), 2030
*Società Italiana Petrolia 2331
*Société d'études chimiques pour l'industrie, 2236
*Société d'études pour la fabrication et l'emploi des engrais chimiques 2059
*Soderberg V. L., 2059
Sotus V., 2189
Solana G., 2200
Soldatov I. T., 2275
Sommer F., 2254
Sommer H., 2295
Sontogyl, M., 2165, 2180
Soper, F. G., 2122
Sorge, G., 2190
Sorrentino, E., 2078
Sosa, E. S., 2177
Spadafora, L., 2150
Späth, E., 2148
Spanner, P., 2191
*Späth, M., 2273
Spasoff, A., 2134
*Speckman, J. B., 2206
*Spears, E. B., 2111
*Spencer D. A., 2065
Spencer, G., 2144
Spencer, M., 2031
Spencer W. R., 2260
*Spicers Ltd., 2289
Spiegler R., 2166
Spiers, C. H., 2321
Spillman, W. J., 2230
Spore, P., 2197
Spring H. W., 2049
Squire, E. J., 2024
Stähler, H., 2103
*Staeger A., 2284
Staeger S. A., 2285
Stäger, H., 2094
Stahl W., 2084
*Stannolabnik Burgdorf A.-G., 2311
Staskaya, D. I., 2197
Stamm, E., 2192
*Stampe G. K. E. H., 2216
*Standard Brands, Inc., 2209
*Staegler, H. G., 2155
Statler, P., 2322
*Staud C. J., 2288
Staudinger, H., 2187
*Steckborn Kanst.-edfa A.-G., 2288
Steed A. H., 2275
Steger, A., 2315
*Steinmann H., 2028
*Stein, P., 2103
Steinberg, S., 2006
Steinchen 2317
Steinkamp, J. H., 2268
*Steinmann W., 2318
*Steinmüller, L. & Co., 2029
von Steinwehls H., 2031
Stempel B., 2208, 2308
*Stenfor F. I. E., 2028
*Stengl R. J., 2059
*Stenkl K., 2137
*Stephan R., 2194
Stephens J. G., 2046
*Stephenson, H. P., 2250
Sterkin E. J., 2184
Stern Ralser, R., 2103
Stettbacher A., 2293
Stende M., 2285
*Stener E., 2028
*Steven, O. L., 2327
Stevens P. W., 2293
Stewardson, D. M., 2238
Stewart, M. J., 2189
Steyer, H., 2036
Stikany, E., 2150, 2325
Stiebel, P., 2063
Stier, T. J. B., 2026
*Still, C., 2275
*Still C. (Firma), 2079, 2272
*Stine C. M. A., 2294
Stis, J., 2208
Stock, A., 2111
*Stock, P., 2315
*Stocking G. W., 2247
*Stockelbach, P. F., 2154
*Stockstrom, A., 2030
*Stöckly, J. J., 2290
*Stokke, R. A., 2106
Stollé, R., 2126
Storsh, H. H., 2247
Strachen, A. S., 2201
Strachan J., 2296
*Strack, E., 2175
Strickhuber, P., 2270
*Strub, W., 2234
*Stretford W. M., 2279
Straumann M., 2080
Strauss A., 2185
*Strebelwerk G. m. b. H., 2029
Strabinger, R., 2074
Streich, H., 2328
*Streich, M. (Firma), 2029
*Streich, H., 2328
*Streich, P., 2178
*Streich, K., 2087
*Strickler, P. F., 2246
Strobel, E., 2059
Strohmenger, A. P., 2111
Stromer C. J., 2301
Stuchovskaya S., 2039
Studerhaus, H. E., 2057
Stuhr J., 2104
*Sullivan, P. W. Jr., 2281
Sulman, H. L., 2011
Sulser, G., 2094
*Summers, S. L., 2245, 2246
Surlany G., 2162
*Sussek, L., 2153
Sutton H., 2103
Suzuki, H., 2195
*Svein, K., 2292
Swanson R. W., 2169
Swietoslowski, W., 2034
*Swingspout Measurs Co., 2256
Symes, E. L., 2319
*Symonds C. W., 2027
Szabá E., 2026
Tachibana, S., 2074
Taffel, A., 2315
Tainter, M. L., 2193
Takahashi, E., 2085
Takahashi, T., 2202

- *Takemura, K. 2311
 *Talisferro, R. H., 2223
 Talmud, D. 2039
 Tama, C., 2033
 Tamatu, S., 2013
 Tammann, G. 2163
 Tanaka, R., 2277
 Tananave, I. 2070
 Tananara, N. A., 2073, 2075
 Tauxery Waste Disposal Committee of Pennsylvania 2228
 Tarlow, P. 2093
 Tate, P. O. II., 2239, 2280
 *Tata, W. H., 2280
 Tausch, J., 2053
 Taylor, F. R., 2131
 Taylor, P. M. II., 2130,
 *Taylor, G. P., 2110
 Taylor, G. L., 2188
 Taylor, J. H., 2046
 Taylor, K. A., 2123
 Taylor, T. W. J., 2013
 *Taylor, W. J., 2304
 Tebl, T. II., 2173
 Teeter, C. B., Jr., 2043
 Teitel, D., 2192
 *Teikunkon Ges. für Drahtlose Telegraphie, 2058
 *Tengler, J., 2254
 *Teppepa, J., 2333
 Tetras, H., 2271
 Terry, H., 2042
 *Terry, J. T., 2109
 *Thas, M., 2311,
 *Theis, F. R., 2323
 Thierault, P. J., 2218
 Thews, E. H., 2101,
 Thiel, A., 2089
 Thiele, H., 2163
 Thies, W., 2171
 *Thilo, E., 2251,
 Thielthwaite, F., 2242
 *Thomas, H., 2061
 *Thomas, A. O., 2253
 Thomas, A. W., 2031,
 *Thomas, H., 2279
 Thomas, W. H., 2278
 *Thompson, C. II., 2254
 *Thompson, W. D., 2209
 Thoms, H., 2215
 Thoreau, J., 2079,
 *Thorp, L., 2245
 *Thoresell, T., 2242
 Thuan, U. J., 2322
 *Thurm, R., 2030,
 Ti, S. P., 2116
 Tidswell, F. V., 2263
 Tildard, D. L., 2308
 Tilmann, J., 2164
 Timmermans, J., 2036
 Tippmann, P., 2260
 Tisdall, P. F., 2174
 Tishchenko, V. V., 2114
 Tiza, I. 2035
 *Toch, M., 2263
 Todd, P., 2040
 *Toda, P., 2010 2061
 *Tolera Reala Mfg. Co., 2313
 Tomaszek, R., 2057
 Tomaszewski, J., 2067, 2098
 Tomiak, G., 2071
 Tommas, G., 2230,
 Tomula, P. S., 2204
 Tonolo, C., 2240
 *Topping, T., 2273
 Torney, H. J., 2031
 Toru, D., 2159
 Townsend, J. R., 2091
 Toyoda, M., 2189
 Trauband, 2240
 Traubenberg, P., 2192
 Traub, H. P., 2207
 *Traube, I., 2218
 Trautschold, R., 2285
 Traut, M., 2034 2036, 2077
 Travers, A., 2068 2170
 *Travers, J. T., 2109
 Treiba, A., 2187
 Treichel, D., 2066
 *Trent Process Corp., 2272
 *Trent, W. J., 2030
 *Trocott, L. C., 2029
 Tula, P., 2187
 Trumble, H. C., 2189
 Trumble, H. M., 2010
 *Truckungs-, Ver-
 schweißungs-, und
 Vergassungs-Ges.
 m. b. H., 2272
 *Trüger, C. H. Jr., 2325
 *Truettner, P. Maschinenfabrik, 2029
 Trogue, C., 2282
 Trost, P., 2120
 *Trotter, W., 2252
 Troadot, H., 2216
 *Trombridge, M. L., 2209 2210
 Truchet, R., 2112
 Truscott, B. J., 2031,
 Truskowski, R., 2154
 *Tuchener, P., 2304
 *Tuchner, H., 2151
 Tsch. A., 2277,
 Tsumaki, T., 2145
 Tsurumi, S., 2077
 Tucker, W. A., 2109
 Tulichunkaja, K., 2161,
 *Tucher, 2262
 Tutom, P., 2172
 Twardy, B., 2269
 *Twiss, D. P., 2321,
 Tyler, P. M., 2218
 Tyutyunnikov, B., 2217,
 Ubalini, I., 2270, 2271,
 Ulrich, H. L., 2219,
 Ulrich, W., 2178
 Umbach, H., 2093,
 Umemo, M., 2161
 Umehiro, Y., 2277
 *Unionchemie belge,
 R. A., 2185
 *Union chimique belge,
 Soc. anon. 2250
 *Unna, Z., 2193
 Unsell, G. P., 2045
 Uota, H., 2242
 *Uphol, P., 2248
 *Urbach, Jr., 2209
 *Urschler, V., 2119
 *Usher, R., 2022
 Uyetsuki, C., 2071
 *Vaders, L., 2108
 Vall, J. G., 2117
 Valenta, J., 2091
 Valentini, W., 2085
 Valk, B. H. van K., 2278
 *Vandyske, H. H., 2181
 Vary, R. A., 2081
 Veach, H. T. Jr., 2219
 Vacker, G., 2187
 van der Veen, A. L. W. H., 2081,
 Velde, H., 2037
 *Verein für chemische und metallurgische Produktion, 2237
 *Vereinigte Aluminiumwerke, A.-G., 2060
 *Vereinigte Drahtwerke A. G., (Dräht-
 oder Röhren A. A.), 2079
 *Vereinigte Stahlwerke A. G., 2108
 Veron, Z., 2024
 Verhoeck, P., 2071
 Verma, M., 2192,
 Vermeulen, D., 2050
 Verchovsky, T. T. II., 2041
 Viale, G., 2182 2209
 Viefelot, J., 2097
 du Vigneau, V., 2116
 Villars, D. B., 2092
 Villaverde, G. V., 2215
 da Vihers, F. J., 2211
 Vilamier, G., 2228
 Vincent, 2221
 Vogt, P., 2073
 Volzner, K., 2271
 Volzner, R., 2177
 Vies, B. J., 2288
 Vilier, B., 2118
 *Vokes, C. G., 2076
 Volmer, I., 2124
 Volodin, A. H., 2181
 *Vorchers, V., 2109
 Voronov, A. I., 2277,
 vom Wark, A., 2283
 Wada, S., 2170
 *Waddell, M., 2305
 *Wadman, H. A., 2256
 Wächterowitz, A., 2116
 Waelch, H., 2214
 Warner, H., 2279
 Wagner, A., 2215
 *Wagner, A., 2106
 Wagner, P., 2237
 *Wagner, H., 2209
 Wagner, H., 2227
 Walawicht, J. H., 2268
 Walls, R., 2201, 2202
 *Walston, J. Corp., 2301
 *Walker, J. H., 2029
 *Walker, J. F., 2279
 Walker, O. J., 2077
 Walker, T. K., 2110
 *Wallace, C. C., 2263
 Wallace, H. H., 2025
 Wallara, C. P., 2026
 Wallen, Laurence Z., 2181
 Walls, W. A., 2032
 *Walter, P., 2306
 *Walters, H. G., 2209
 Wauk, G., 2121
 Wang, C. C., 2163
 Warne, K., 2119
 *Warren Tool Steel Co., 2227
 Washburn, P. A., 2081
 Washburn, H. W., 2111
 Watanabe, M., 2135
 Waters, R. W., 2201
 *Watkins, W. F., 2106
 *Watson, C. H., 2275, 2278
 *Watson, H. C., 2305
 Wawo Kuentor, P. K., 2192
 *Weber, C. B., 2286
 *Weber & Co., 2260
 *Weber, J., 2162
 Wedekind, F., 2072
 Webster, H. C., 2049
 Weichardt, W., 2178
 *Weidert, P., 2259
 Weigel, P., 2083
 *Weiland, H. J., 2187,
 Weinberger, W., 2039
 Weiss, J., 2089
 Weiss, J. J., 2044
 Weisbach, K., 2134
 Weisberger, G., 2267
 Weisfeld, W., 2266
 Weist, W., 2051
 *Weiskopf, P. A., 2087
 *Weitsch, D., 2236
 Weisland, H. R., 2221
 *Weitzel, W., 2296
 Wenner, W. P., 2197
 *Werke, P. A.-G., 2273
 *Werner, K., 2288
 Werner, W., 2270
 *Werschen - Weissen-
 stein Braunkohlen,
 A.-G., 2213
 Wess, W., 2074
 *Wesche, H., 2274
 *West, C., 2209
 West, C. J., 2285, 2287,
 West, P. D., 2218,
 Westbrook, F. A., 2274,

- Westby C N., 2248
 Weston, W A R D., 2233
 *Wetherbee, A. U., 2027
 Whedbee, E., 2221.
 Wheeler, R. V., 2271, 2293
 White, D R., 2065
 White, H. A., 2084.
 White, P B., 2188.
 *Whiteley, J T., 2109
 *Whittemsey, T., 2332.
 Wiberg E., 2066
 *Widerkehr, P., 2030
 Wiklake, H. C., 2267
 *Wiedbrauck, E., 2133
 Wiegand, W B., 2277
 Wieland, H., 2131
 *Wiener, J A., 2221
 Wiemann H., 2229, 2230
 *Wietzel, R., 2231
 Wightman, E P., 2065
 van Wijk, D J R., 2207
 Wilcoxon, F., 2169
 de Wild, A. M., 2307
 *Wilde, D F., 2239
 Wiley, R. C., 2071
 Wilhelm, J D., 2051, 2098, 2099
 Wilhelm, E., 2248
 *Wilkinson, E W., 2242
 Willard, M. L., 2025
 Willets, W R., 2252
 *Williams E. R., 2103, 2107
 Williams, J B., 2073
 *Williams, W D., 2263
 *Williams W H., 2153, 2154
 *Willshaw R., 2332
 *Wilson, E. G., 2375.
 Wilson, F J., 2122
 Wilson, J A., 2324
 Wilson O., 2710
 *Wilson, T. D., 2275
 *Winder, F J., 2106
 Wisegradec, H M., 2162
 *Wishup, W., 2029
 *Wintermute, H A., 2061
 Witz, P., 2121
 *Wlachos R. A., 2181
 Wisler, W A., 2245
 Wither, W O. M., 2093
 Witzemann, E J., 2114 2118.
 Wolbing H., 2031
 Wokes, F., 2741
 Wolf, H., 2294
 *Wolf, R. B., 2291.
 Wolff, H., 2307, 2309
 Wolff, P., 2245.
 Wolff W., 2164
 Wolfe, J B., 2201
 *Wolfson, S L., 2291
 Wollenberg H., 2023
 *Wolser, P., 2259
 *Wolser, R., 2739
 Walter, K., 2045
 Woo, C., 2072
 Woolf, J A., 2082.
 Wood, A. A., 2165
 Wood, D R., 2205
 Woodard, H. Q., 2163
 *Woodford, W. H., 2294
 Wooster, N., 2035
 Wormam, R., 2234
 v. Wraugell, M., 2229
 Woods, P., 2175, 2023
 *Wright, D K., 2042
 *Wright, E C., 2109
 *Wright, E H., 2209
 Wright, G P., 2189
 Wright, L. H., 2267
 Wright, R., 2040
 Wrighton, H., 2069
 Worth, K., 2309
 Wastrow W., 2111
 *Wyatt K. M., 2259
 *Wylie, J H., 2254
 Wyman, J., Jr., 2162
 *Yacco, S. A. F., 2281
 Yamazaki, T., 2330
 Yarwood, C., 2233.
 Yashchenko V., 2217
 *Yates, C. E., 2059
 Yatsavitch, M O., 2096.
 *Yensen, T D., 2107
 Yokoyama, H., 2207
 Yoshie, S., 2231
 Yoshimura J., 2078
 Yoshoka T., 2160
 Yost D M., 2072
 *Young P W., 2216
 Young, H C 2235
 Yur'ev, K., 2058
 Zablocki, B., 2190
 Zabev, D., 2031
 Zacharewicz, E., 2243
 Zahn, 2031
 *Zahn & Co., D m. b. H., 2029
 Zakharov, E R., 2079
 Zambalos, P G., 2125.
 Zechmeister, L., 2172
 von Zeerleder, A., 2087
 Zendenfeld, S., 2050
 Zeile, K., 2157.
 *Zeiss, C. (Firma), 2053.
 *Zentus, D G., 2026
 Zeleny, L., 2117.
 Zelinski, N D., 2044
 *Zellstoffabrik Waldhof, 2258
 Zeman, J., 2264.
 *Zerekudze E F., 2111
 Zervas, L., 2151
 Zetsmbe, P., 2116
 Zeynek, R., 2214
 *Ziegler, K., 2154
 Ziegler, N. A., 2107
 Zibermiotz, V., 2078
 Zimmermann, A., 2294
 Zink, R., 2034
 Zlataroff, A., 2237
 Zmaczynski, A., 2034
 Zondck, B., 2175
 Zondek, S D., 2176
 Zoellner, E. A., 2126
 Zumbach, P. H., 2256
 Zuchner, H B., 2050
 *Zumbach - Werke Kaiserlautern, A D., 2028
 Zumbach, J., 2192
 *Zwicky, H., 2029

CHEMICAL ABSTRACTS

Editor: E. J. Crane, Ohio State University, Columbus, Ohio.
Associate Editors: Elmer Hockett, Helen Gama and Leonard T. Capell
Assistant Editors

M. S. Anderson	O. J. Demarest	P. B. Hawk	L. A. Olney	J. J. Skinner
W. L. Badger	Alden H. Emery	P. H. Howe	A. M. Patterson	W. R. Stearns
C. E. Barton	W. O. Emery	C. H. Kerr	T. G. Phillips	C. O. Storm
Edward Barrow	W. F. Fargher	H. A. Lepper	A. N. Richards	H. M. Symmes
S. R. Benedict	A. C. Fieldner	R. K. Marshall, Jr.	Richard Rimbach	P. F. Underhill
F. C. Black	C. G. Fink	A. R. Middleton	Allen Rogers	H. G. Wells
F. C. Brown	C. N. Frey	Harlan S. Mizer	C. A. Rouillet	C. J. West
C. E. Curran	H. W. Gillett	Chas. B. Morrey	A. H. Sablin	H. T. Wherry
J. K. Dale	R. A. Gortner	C. R. Monroe	J. P. Schaefer	H. P. Wightman
C. C. Davis	W. T. Hall	W. A. Noyes Jr.	K. Scherubel	J. C. Witt

CONTENTS

Patents	i	14. The Fermentation Industries	2515
Author Index	ii	15. Pharmaceutical Chemistry	2517
1. Apparatus and Plant Equipment	2113	16. Acids, Alkalies, Salts and Sun-	2521
2. General and Physical Chemistry	2113	drives	
3. Subatomic Phenomena and Radio-		17. Glass, Clay Products, Refractories	2524
chemistry	2117	and Enamels Metals	
4. Electrochemistry	2123	18. Cement and Other Building Ma-	2527
5. Photography	2178	terials	
6. Inorganic Chemistry	2180	19. Fuels, Gas, Tar and Coke	2541
7. Analytical Chemistry	2185	20. Petroleum, Lubricants, Asphalt	
8. Mineralogical and Geological		and Wood Products	2551
Chemistry	2190	21. Cellulose and Paper	2553
9. Metallurgy and Metallography	2192	22. Explosives and Explosions	2555
10. Organic Chemistry	2411	23. Dyes and Textile Chemistry	2570
11. Biological Chemistry	2423	24. Paints, Varnishes and Resins	2573
12. Foods	2430	25. Fats, Fatty Oils, Waxes and Soaps	2582
13. General Industrial Chemistry	2444	26. Sugar, Starch and Gums	2584
14. Water, Sewage and Sanitation	2448	27. Leather and Oils	2587
15. Soils, Fertilizers and Agricultural		28. Rubber and Allied Substances	2591
Poisons	2504		

PATENTS

A statement giving information as to how to obtain patent specifications and drawings, both United States and foreign, is to be found on page i of *Chemical Abstracts*, 24, No. 5

Published by THE AMERICAN CHEMICAL SOCIETY, which also publishes *The Journal of the American Chemical Society* and *Industrial and Engineering Chemistry*

Communications for the editor should be addressed to E. J. Crane, Ohio State University, Columbus, Ohio.

Remittances for annual dues (\$15.00) in the American Chemical Society or for subscriptions to this *Journal* (\$7.50 per year), changes of address, orders for subscriptions or back numbers, claims for missing numbers and all matters intended for the *Proceedings* should be sent to the Secretary, Charles L. Parsons, Mills Bldg., Washington, D. C. Foreign postage is extra, being \$4.60 per year added to the dues of members residing outside the United States, its possessions and the countries of the Pan American Union and \$1.50 per year for *Chemical Abstracts* alone, except that only one-half of these amounts is added for journals going to Canada. The countries of the Pan American Union are Argentina, Bolivia, Brazil, Chile, Colombia, Costa Rica, Cuba, Dominican Republic, Ecuador, Guatemala, Mexico, Nicaragua, Paraguay, Peru, Republic of Honduras, Republic of Panama, Salvador, Spain and Uruguay.

The Council has voted that no claims will be allowed for copies of journals lost in the mails unless such claims are received within sixty days of the date of issue, and no claims will be allowed for issues lost as a result of insufficient notice of change of address. "Missing from files" cannot be accepted as the reason for honoring a claim—Charles L. Parsons, Secretary.

Published by the American Chemical Society, Easton, Pa. Secretary's Office: Mills Bldg., Washington, D. C. Entered as second-class matter at the Post Office at Easton, Pa. Acceptance for mailing at special rate of postage provided for in Section 1103, Act of November 3, 1917, authorized July 13, 1918. Mack Printing Company, Easton, Pa.

AUTHOR INDEX

[An asterisk (*) indicates that the reference is to an abstract of a patent.]

- Ahlqvist F, 2401
 Alarshakien T, 2417
 2443 2463 2493
 Alkal F, 2410
 Alkali I, 2493
 Alramson H, 2451
 Addestone J A, 2343
 Adino G F, 2111
 Adova A, 2444
 Agar C C, 2502
 Akademischen Verein
 flütte E, 2496
 *Aktiebolaget Ljung
 ströms Ångturbin,
 2500
 *Aktiebolaget Sepsa
 tor 2333
 *A G Brown, Boven
 & Cae 2336
 Akulov, N S, 2310
 Alberti, E, 2453
 *Alberta W, 2409
 Alfarretti C, 2490
 Albright F, 2469
 *Altehande, A.
 (Firma), 2374
 *Alexander, C. M.,
 2337
 Alhous A, 2544
 Allen F L, 2353
 *Allgemeine Elektrizi-
 tät-Ges., 2577
 *Allgemeine Ges für
 Chemische Indus-
 trie 2537
 *Althott E A, 2335
 *Altus, N, 2374
 Altnberg, N, 2342
 Altköder, F, 2373
 Altpeter J, 2435
 Amburg, C R, 2535
 Ambler, J A, 2545
 Amerink, J P H, 2517
 *American Potash &
 Chemical Corp.,
 2329
 *Anderson D & Son,
 Ltd., 2554
 Anderson, H H, 2453
 *Anderson K C, 2277
 Andena, K., 2342
 Andrew J H, 2401
 *Andrews C W, 2500
 Andrews E, 2475
 *Anst, F, 2540
 Angeletti, A., 2425
 Ansel, A., 2434
 Anselmi, A., 2554
 Antuneso G, 2478
 Appleton, E V, 2333
 Appleton J L T., Jr.,
 2453
 Appleyard, K C, 2544
 *Aps, J E, 2494
 Arakawa, S., 2454
 Artukow A, 2414
 Artukow A, 2414
 Arens, H, 2378
 Armandet A, 2470
 Arndt P, 2427
 Arndt K, 2394
 Arnot J M, 2509
 *Art 2450
 *Ashcroft E A., 2409
 Asher L, 2463
 Ashley S E, 2366
 Askew, P., 2379
 Asperger, H., 2473
 Anthony, W T., 2349
 *d Artack E L., 2369
 Atsuki K., 2500
 Atbel E, 2455
 *Auer L., 2554
 Aumeras M, 2354
 Aunis L, 2509
 *Autogenwerk Sonne
 G. m. b. H., 2335
 Averil H P, 2493
 *Avronomov V S.,
 2500
 *Azeo S P., 2378
 *Atell J M, 2336
 Ayar C V R., 2503
 Ayar, N K., 2464
 Bach H., 2501
 Back, R, 2374
 Bahla W E, 2334
 Baby H., 2411
 Bailey, E. H. S., 2390
 *Bailey, G B., 2333
 *Bailey, G C., 2509
 *Baker, G B., 2336
 *Baker J F, 2334
 Baker, R., 2372
 *Baker T T., 2378
 Baker W J., 2524
 *Balachowsky, D.,
 2537
 Balinski, I A., 2360
 *Baint I., 2408 2429
 Ball, G H., 2505
 Ball, M., 2401
 *Balfet, D., 2335
 Palmer G, 2424
 *Bomphylde, J W.,
 2410
 Bancroft, W D, 2459
 *Bandus F., 2537
 Banerji, S N., 2349
 *Banister A., 2531
 Banister, S H., 2424
 Banister, W J., 2411
 Banov A V., 2366
 Banushal, E J., 2474
 *Barbet, E A., 2317
 Barstow, G A., 2394
 Barbat, M., 2513
 Barclay, W K., 2403
 Barclay, G., 2435
 Barker R E., 2393
 Barua, I., 2554
 *Barward D P., 2506
 Baroda, D., 2366
 *Barret R S., 2554
 Barrenschon H K.,
 2471
 Barrow, E S G., 2475,
 2492
 Bartholomew, E, 2553
 Bartrum, J A., 2392
 *Bashlov I V, 2530
 Bates, L P., 2357
 Batetree T., 2343
 Bauer K H., 2518
 Bauer K., 2475
 Bauer, W., 2469
 *Beuer W, 2551
 Baumann E J., 2453
 Baumman E V H.,
 2532
 Baver L D, 2505
 Bay Z, 2362
 Bazurina, E N, 2457
 Beall J N, 2532
 Beaufoer H A., 2452
 Beaver D J., 2591
 Beck W, 2403
 *Becker C E., 2408
 Becker J E., 2463,
 2484
 *Beckmann H., 2597
 *Beckel, Ltd., 2494
 Beck P A. A. van der,
 2306
 Begbie, R S., 2453
 Behrens B., 2453
 Bell, M A., 2471
 Benedicks C., 2399
 *Bennett L C V., 2406
 Benjama H R., 2465
 *Benner R C., 2531
 Benton F T., 2544
 van den Berg G., 2555
 Berger, E., 2443
 von Bergkamp, E S.,
 2394
 Bergmann E., 2428
 *Bergner O., 2404
 Berl E., 2559
 *Berl, F., 2334
 Berlin L E., 2504
 Berre E., 2522
 Berwick W., 2334
 Best C H., 2450
 *de Béhne, G S F.,
 2334
 Betz A., 2333
 Beutel E., 2579
 Beverdam H B., 2503
 Beyer A., 2571
 *Beard F., 2540
 Bhargava, S, 2359
 Bhattacharya S S., 2345
 Bicking G W., 2512
 2563
 Biersch, R., 2447
 *Bilings, H J., 2533
 *Bilster J., 2375
 Bilson, 2397
 Billa, C E., 2469
 Biltz W., 2542
 Bissel L., 2470
 *Birchall T., 2437
 Birke J., 2361
 Birstein G., 2354
 Bissas, H., 2426
 Bittzer, E C., 2400
 Blacklock L., 2506
 Blackstone, A., 2470
 Blackman, J., 2542
 Blaser, B., 2332
 Bicker B., 2460
 *Bliss A V., 2550
 Bloedel M F., 2392
 Bloomfield A L., 2469
 *Blumner E., 2495
 Blumfeld J., 2500
 Blumgart H L., 2452
 Bode H., 2354
 Bodenstein M., 2396
 Bodnar, J., 2451
 Boeckner C., 2360
 *Boehme H T A G,
 2531
 *Boehringer C F &
 Soehne G. m. b. H.
 2441
 Boekholt K., 2511
 *de Boer J H., 2525
 *Boha, M., 2336
 Bodin A R., 2452
 Bodin W B., 2333
 *Boehm, C., 2378
 Bondy, H P., 2592
 Bone W A., 2542
 2569
 Borchardt E., 2469
 Borden J., 2453
 *Borow, V L., 2335
 Borowman J A., 2521
 Boros A., 2518
 *Borag A. G. m. b. H.,
 2405
 *Borzykowski, B.,
 2578
 *von Bove J., 2375
 Botset, H G., 2346
 Bouchard, J., 2367
 Boulienne, M., 2450
 2461
 Boulienne, R., 2460,
 2461
 Boutsic A., 2367
 *Bowers D W., 2411
 Boxer H., 2371
 Boyden R E., 2465
 Boye E., 2344
 *Boykin R O., 2335
 *Boyne H E., 2337
 *Bozel Maletta (Soc
 industrielle de pro-
 duits chim.) 2508
 Braak, H R., 2503
 *Brace, P H., 2376
 *Brackelsberg C., 2404
 *Bradley, H L., 2379
 *Bradley, L., 2567,
 2568
 *Bradshaw G D.,
 2333, 2337
 Brady, G S., 2496
 Brambilla A., 2425
 *Brancat V., 2506,
 2537
 *Brand M., 2497
 *Bransky A. H., 2409
 *Brasert, H A., 2550
 *Brass J., 2429
 Brann K., 2411
 Braun W., 2490
 Braune H., 2554
 Bravo O A., 2547
 Breckner R., 2412
 Breddin, 2518
 Brechtoldt W., 2400
 Brewer, F M., 2390
 Breckwedde, P G.,
 2334
 Brigg G E., 2450
 Briggs, K A., 2415
 Brinkman, H C., 2359
 Brinkman, K., 2504
 Brinsinger, H., 2347
 2350 2351
 Brinsinger W., 2347,
 2350, 2351
 Brinze, H T., 2390
 *Brinze F J., 2429
 *British Celanese Ltd.,
 2566 2567 2578
 *British & Dominion
 Fertilizer, Ltd., 2410
 *British Dyestuffs
 Corp Ltd., 2573
 *British Glue &
 Chemicals Ltd,
 2590
 Brutsche, E., 2526
 *Broemel L E., 2409
 Brown, H R., 2509
 Brown, J C., 2391
 Brown, M A., 2462
 *Brown, R., 2337
 Brown, R H., 2404
 Brüning H., 2344
 *van der Brugge, A.,
 2532
 de Brun, T L., 2360
 *Brun L C., 2567
 Brunel B J., 2541
 Brunsen, H C., 2372
 *Brunson H A., 2537
 Brous J H., 2551
 Bruzas (Sme.), 2415
 Bruts, B., 2369
 *Bryukhov, A. A.,
 2540
 *Buchanan, G H.,
 2515
 Bucherer, H T., 2346

- *Buchhaas, E., 2504
 *Buchholz, M., 2377
 Buckley, T. A., 2359
 Buddie, H., 2439
 Budnikov, P. P., 2336
 *Burger, C. H., 2556
 Buerger, M. J., 2341
 Birstenbinder, R., 2379
 *Büsching, W., 2527
 Buxington, R. M., 2340
 Bunce, E. H., 2520
 *Bunge, C., 2531
 Bunte, K., 2516
 *Burch, C. R., 2376
 *Burdick, I. C., 2532
 Burke, S. P., 2542
 Burkert, A. W., 2533
 *Burkholder, T. M., 2339
 Burr, M. L., 2483
 Byall, S., 2535
 *Byland, H. C., 2540
 Bytchier, A., 2570
 Cady, H., 2390
 Cagliotti, V., 2335
 Caprasso, A., 2383
 Cailliet, O. R., 2446
 *Caire, P., 2517
 *Calcutt, W. S., 2507
 Calhoun, J. A., 2453
 Calvi, G., 2499
 Cambr, L., 2393
 *Cambren, A., 2595
 Campbell, F. C., 2489
 Campbell, J. S., 2360
 Campbell, W. B., 2564
 Canales, T. M., 2385
 Castacuerne, J., 2478
 *Cappel, H. C., 2407
 Carlson, C. G., 2532
 *Carlson, C. L., 2537
 Carlson, H. A., 2474
 Carnahan, G. H., 2501
 *Caro, N., 2496, 2498, 2527
 Carobbi, G., 2390
 *Carothers, J. N., 2436
 Carpenter, C. H., 2532
 *Carpenter, J. B., 2531
 Carr, C. J., 2446
 Carrelli, A., 2364
 Carner, W. H., 2503
 *Carrington, F. G., 2408
 *Carson, C. M., 2507
 Carson, F. T., 2465
 *Carveth, H. R., 2530
 *Casale-Sacchi, M., 2436
 Casberg, C. H., 2395
 Cassar, H. A., 2333
 Cassel, H., 2344
 Castiglioni, A., 2534
 Castille, A., 2364
 Catalan, M. A., 2361
 Catalano, V., 2466
 Canquill, G. (Mlle.), 2420
 Cazran, U., 2388
 Celis, S. A., 2478
 Cesbron, R., 2460
 Challenger, F., 2516
 Chalmers, A., 2522
 Chamberlain, J. M., 2535
 *Chance Bros & Co., Ltd., 2536, 2537
 Chapin, W. H., 2356
 *Chastot, J. C., 2533
 Chaudron, G., 2404
 Chavanne, G., 2420
 *Chemiefabrik, G. m. b. H., 2544
 *Chemische Fabrik Joachim Wernik & Co., A. G., 2436
 *Chemische Fabrik vorm. Sandoz, 2572
 Chesnokov, V. A., 2437
 Chichabalin, A. B., 2531
 Chidester, F. F., 2463
 *Child, R. O., 2554
 Child, W. H. J., 2361
 *Chinson, Ltd., 2514
 *Chinois Gydyzser & Vegyészeti Termék Gyára R. T. (Kereskedési és Váltó) 2515
 *Chitty, C. W., 2193
 Chulmky, L., 2520
 Christiansen, M., 2562
 Chrusle, W. A., 2391
 Chulmky, L., 2520
 Chocchi, L., 2563
 *Clapp, A. L., 2569
 Clark, H. W., 2501
 Clark, N., 2565
 *Clark, T. W. P., 2515
 Clarke, G., 2453
 Clarke, T., 2492
 Clarou, J., 2570
 Claude, P. E., 2473
 Claus, W., 2403
 Clausmann, 2397
 *Claxton, L., 2541
 Clayton, D., 2402
 Clayton, R. E., 2514
 Clegg, R. L. G., 2391
 *Clifford, A. M., 2597
 Clifford, W. M., 2461
 Clove, P. D., 2493
 Clotworthy, H. R. S., 2560
 Clowes, F., 2390
 *Clutterback, A. H., 2336
 Cock, S. A., 2513
 Cockrill, J. R., 2469
 Codegone, C., 2533
 *Collett, O. T., 2407
 *Colley, S., 2437
 Cohn, W. M., 2531
 Coleman, J. H., 2390
 Culey, C., 2368
 *Coley, H. F., 2407
 Colla, S., 2458
 Colless, W. S., 2453
 *Collins & Co., 2551
 Collins, C., 2391
 *Collinge, W. R., 2529
 *Collier, F., 2533
 Coltof, W., 2570
 *Combes, P., 2494
 Combesco, D., 2478
 *Compagnie française pour l'exploitation des procédés Thomson Houston, 2582, 2584
 *Compagnie générale des conduites d'eau, Soc. anon., 2408
 *Compagnie Internationale pour la fabrication des caisses et pétroles, 2338
 *Compagnie des Surbauffeurs, 2549
 *Connors, C. W., 2587
 *Constable, F. H., 2374
 *Consortium für elektrochemische Industrie Ges., 2436, 2438, 2441
 *Cook, A. H., 2494
 *Cook, C. W., 2335
 *Cooley, J. L., 2557
 Cooper, J. R., 2570
 Corbida, A., 2453
 *De Coriolis, P. G., 2336
 *Cornell, F. G. Jr., 2493
 Corradini, G., 2473
 Coste, J. M., 2582
 Coster, H. M., 2525
 Coughlan, R. E., 2501
 *Coulson, B. P., 2334
 Coulson, J. G., 2513
 *Cowan, H. W., 2541
 *Comper Co's, S. H., 2375
 Cox, C. H., 2453
 Cox, L. H., 2336
 Coyle, F. B., 2402
 *Crabbs, G. D., 2541
 *Cracum, E. C., 2474
 *Crax, T. J., 2532
 *Craxton, E., 2464
 *Crennell, J. T., 2372
 *Cripe, S. C., 2556
 *Cripe, P., 2492
 *Cripe, A., 2551
 *Crut, R. H., 2354
 *Crose, R., 2553
 *Cruse, W. V., 2493
 *Cuba, M., 2473
 *Cupke, Z., 2519
 *Cuffe, F. W., 2376
 *Cullen, G. E., 2483
 *Culp, P. B., 2465
 *Cummins, A. D., 2578
 *Curphy, T. J., 2474
 *Curry, J. R., 2344
 *Cuthbertson, J. W., 2370
 Dadeau, A., 2364, 2365
 Davies, K., 2394
 *Dakers, J., 2360
 Dalmon, C., 2454
 *Daley, W. A., 2337
 *Dall, T., 2337
 Dählmann, A., 2478
 Damerell, V. R., 2339
 Dammköhler, G., 2350
 Damm, Z., 2458
 *Dantal, L., 2378
 *Darwins Ltd., 2410
 Das, R., 2356
 Datt, S., 2488
 David, N. A., 2483
 *Davis, A. R., 2340
 *Davis, N. R., 2376
 Davis, S. S., 2456
 Day, J. N. F., 2411
 *Dean, H. P., 2544
 *Dearborn, R. J., 2556
 *Debauche, H., 2549
 Debenstedt, K., 2562
 Decker, A. C., 2500
 Decker, R., 2516
 De Cost, P., 2356
 Deforme, J., 2544
 *Dehler, F. C., 2498
 *Delas, F. X. J. A., 2337
 Demolon, A., 2510
 *Dempsie, R. & J., Ltd., 2551
 Denney, C. G., 2531
 Denng, H., 2469
 Derksen, J. C., 2349
 Desch, C. H., 2395
 *Deutsch, H., 2441
 Deutch, P., 2545
 *Deutsche Gold- und Silber-Scheideanstalt vormals Roce
 aler, 2375, 2442, 2529, 2537
 *Deutsche Luftfilter-
 Industrie m. b. H., 2335
 Dewael, A., 2416
 Diellele, P., 2434
 *Diamant, G., 2374
 Dickie, J. K., 2547
 Dietrich, K. R., 2541
 Dietz, P., 2377
 Ditt, H. B., 2463, 2469
 *Dillger, C., 2374
 *Dippon, G., 2455
 Dingemans, L., 2469
 *Dittmer, S., 2531
 Dittmar, R., 2501, 2593
 *Dixon, T. R., 2517
 Doan, P. J., 2491
 Doboryski, D., 2340
 Dodds, H. H., 2584
 Döpel, R., 2357
 Dogon, A., 2451
 Dohm, W., 2335
 *Donat, V. F., 2410
 Donelson, P., 2465
 Dork, M., 2467
 *Douglas, C. B., 2504
 *Douglas, W. A., 2507
 *Doubtitt, P. H., 2491
 *Douty, A., 2409
 *Dow Chemical Co., 2376
 Draghetti, A., 2455
 *Drake, J. L., 2536
 *Drake, T. J., 2526
 *Dreling, V. J., 2590
 Drury, A. J., 2372
 *Drew, R. B., 2521
 *Dreyer, K., 2557
 *Dreyfus, C., 2577
 *Dreyfus, H., 2435, 2457, 2567, 2578
 Dreyfuss, C., 2508
 Driesen, L. A., 2570
 Drosdov, S. S., 2441
 *Drosdov, V. F., 2590
 Drury, A. N., 2464
 Dubois, E., 2547
 *Dubovitz, H., 2337
 *Ducaamp, A. J., 2558
 Durken, H., 2450
 Duvel, C. O. Jr., 2496
 Duffenback, O. S., 2372
 *Duffield, F. L., 2407
 Dugé, G., 2523
 Dujardin, O., 2589
 Duke Elder, F. M., 2481
 Duke Elder, W. S., 2441
 Dulère, W. L., 2470
 Dumanski, A., 2345
 Du Mond, J. W. M., 2359
 Duncan, W. H., 2381
 Dunlap, W. M., 2405
 *Dunlop Rubber Co. Ltd., 2590
 Dunsheath, P., 2373
 Du Pont de Nemours, E. I. & Co., 2437
 Dupré la Tour, F., 2418
 *Dupuy, H., 2338
 Dushman, R. H., 2454
 Dutcher, R. A., 2461, 2462
 Du Toit, M. S., 2505
 Dutrell, J., 2570
 Dvorak, H. J., 2474
 *Dyckerhoff, E., 2533
 Dyer, H. M., 2481

- Gudris, N. 2358
Güdenst, E. M., 2353.
Güntherhulze, A., 2372
Günzberg, A. M., 2359
Guerin, E., 2355, 2357
Guerril, G. 2150
Guertler, A. 2102
Guichard 2397
Guillet L. 2403
Gums W. 2493
Gupta S. 2360
*Gustin, D. S., 2377
*Gutehoffnungshütte
Oberhausen A.-G.
2543
Guthmann H. 2444
Gvodieva, O. M., 2341
Gwathmey J. T. 2494
Gwozd, J. 2544
*Gyles, T. B., 2404
Haas P. 2155
Haas P. 2179
Haaber P. 2157
Haaberlandt L. 2469
Habs H., 2143
Hadfield C. F., 2323
*Hadfield R. A., 2350
Hadow H. J., 2351
*Hägglund E., 2545
*Häger, A. L., 2333
*Häger, L. B., 2333
Hagman, S. M., 2591
Hahn, P. L., 2333
*Hailwood, A. J., 2573
Hale, F. E., 2149
Hall, A. J., 2371
Halla, F., 2590
Halpe, O. M., 2562
Halverson, J. G., 2462
*Hamburg, M., 2117
Hamid M. A., 2356
*Hamil, J., 2597
Hampil, B., 2454
*Hansen, C. H., 2373
Hansen, P., 2500, 2502
Hanson, D., 2396
Hanslik, P. J., 2446
Happel, J., 2356
Happer J. R., 2564
Hard, E. W., 2532
Harder, O. H., 2533
Hardie, T., 2548
Harker, C., 2366
Harkins W. D., 2315
Harris, L. J., 2164
*Harrison, C. F. R., 2508
Harrison T. R., 2493
Harrop, G. A., 2475
Harrow, B., 2356, 2470
*Harslow Chemical Co., 2527
*Harter, I., 2337
*Hartley, E., 2504
Hartley, H., 2351
*Hartmann, E., 2337
*Hartmann, M., 2337
Hartwell, A. Jr., 2397.
*Harvel Corp. Tbc., 2584
Harvey, W. E., 2401
Haselhoff P. 2509
Hashi, K., 2542
Haskins, H. D., 2453
Hata, C., 2424
*Hatakeyama, H., 2525
Hatcher, W. H., 2351
Hatzmaker, P. 2372.
Hatori, H., 2419
Haun, P., 2509
Hauptstein, P., 2472
Hauser, P. A. 2551
Hanslatter, P. L. 2501
Havestadt L. 2145
*Hawenta - Platten Ges. m. b. H. 2131
Hawkinson A. 2140
Hayden C. C. 2162
Hayes A. 2101
Hayward A. M. 2150
Hesley M. A. 2375
*Healy J. A. 2411
*Heath G. F. 2335
*Heibel H. 2419
Heidfeld K. 2163
*Heidlund W. T. 2199
Heike W. 2100
Heibrunn L. A. 2474
Heim de Balaz 2466
Heim G. 2190
Hense K. 2579
Henta K. 2512
Hektoren J. 2474
*Heller O. 2549
*Hene C. 2524
Heming J. C. 2491
Hepturn, J. S. 2521
Héray A. 2319
*Hebert P. G., 2408
*Herczeg H. 2525
Herd C. E. 2399
Herrmann C. 2527
*Herrli A. G. 2551
*Herrmann, W. O., 2441
Hertel P. 2311 2332
Hersfeld F. 2143
Hess, A. P., 2143
Hess, K. 2319
Hess W. C., 2119 2470
*Hessen B., 2357
*Hetherrington, H. C. 2514
*Heuser, H., 2517
Heymann, P., 2164
Heymanns, A., 2429
Hibbert, H., 2561
Hibbert, L. J., 2374
Hibino T., 2110
Hickmann, E. M., 2154
Hiedemann, E., 2310
Higgins, J. H., 2100
Hill H. B., 2552
Hill, R., 2147, 2445
Hird, J., 2793
Hirsch 2339
Hitz, H., 2496
Hoar, T. P., 2102
Hoch J., 2367
Hochberg B. M., 2342
Hodgson, H. H., 2363
Höfer, E., 2397
Hoffman, 2139
Holman, I. L., 2498
Holland, H. W., 2542
Hollings H., 2543
*Hollösy, Z., 2566
Holmes, A. D., 2163
Holmyard P. J. 2354
*Holt, C. F., 2551
Holter, H., 2447
*Holtmann R., 2356
Holf, J., 2492
*Holzner P. 2336
*Holzverkohlung - Industrie, A. G., 2496
Homerberg V. O., 2397
Honeywell, E. M., 2468
Honeywell H. E., 2461, 2462
Hood, G. R., 2371.
Hoover C. P. 2500
*Horvath, E., 2466
Horvath J., 2514
*Howell M. 2174
*Hoover P. F., 2554
Hotherhall A. W. 2364
*Hottenroth W. 2466
Hough J. P. 2553
Houghton W. P. 2554
*Howard P. A., 2554
*Howard G. P. 2556
Howard J. W. 2391
Howatt J. L., 2512
Hower J. H. 2496
*Hoyt, T. 2315
Hruba J. H. 2100
*Hubbard L. R., 2541
Huckett, H. C., 2513
Hulguin, H. 2500
*Hutmeyer C. 2155
Huestis H. L. 2195
Hüttig G. F. 2145
Hukumoto A. 2167
Humes C. H. 2401
Humphrey S. W. 2334
*Humphreys C. J. 2360
Hun O. (MFe) 2343
Huni P. 2562
*Hungana M. 2164
gygyár Ltd., 2527
Hunziker H. A., 2465
Huster, O. 2177
*Hunsady I. 2510
Huppmann G., 2426
Hurd E. C., 2501
*Hurrell G. C., 2334
*Hussey, I. H., 2564
Husson M., 2540
Huttm, A., 2381
Hyslop, J. P., 2521
Ichibá, A., 2431
Ichihara, K. 2445
Ichise, M., 2550
I. G. Farbemundatne A. G., 2337, 2338 2375, 2379, 2380 2111, 2431 2436, 2437, 2134 2439, 2440 2442, 2152, 2497, 2498, 2114, 2515 2517, 2523 2524, 2524, 2529 2531, 2532 2533 2541, 2549 2549 2550 2556, 2554 2566, 2567, 2572 2573, 2574 2575 2576, 2577, 2578, 2579, 2580 2581, 2584, 2590, 2597, 2599
*Ignatzenko, S. N., 2501
*Ignatovaku, L. P., 2401
*Ignatyev, C. N., 2575
Itaka I., 2153
Itan, B., 2313
*Itiner, R., 2330
Imhoff K., 2591
*Imhoff, K., 2504
Imhoff, W. G., 2403
Immerschutt, J., 2331
*Imperial Chemical Industries Ltd., 2411, 2437, 2442, 2549 2559, 2597, 2598, R. F., 2557.
*Industrial Process Corp., 2594
*Ingersoll, G. B., 2558
*InterCarbo Soc anon pour la carbonisation et le traitement catalytique des combustibles, 2549
*International Patents Development Co. 2544
*International Pre-cipitation Co. 2177
Inoue K., 2393
Inoue A., 2511
Inoue A. A., 2404
Inouen M., 2451
*Iredell C. V. 2107
*Ire M. 2174
*Iscoveco H., 2524
Ishita T., 2542
Ishihara M., 2460
Ishiguro T., 2467
Ihano A., 2151
Ito H., 2159
Isakov, N. N., 2149
Isakura N., 2445
Iwan J., 2141
Jacob K. D., 2404
Jacob, W., 2140
Jacobsen R. C. 2310
*Jack W., 2576
*Jaeger P. W. 2339
Jahr H., 2314
Jakoby L. 2397
*Jalowitz F., 2517
James R. F., 2564
Jander, W., 2341
Jankowska H., 2470
Jelinek J., 2331
Jeltinghaus W., 2556
Jenkins, C. H. M., 2397
Jensen P. C., 2374
Jensen, H. L., 2406
Jewell W. R., 2492
Julin S., 2475
Johaid, N. J., 2359
Johnson, A., 2393, 2547
Johnson F., 2562
*Johnson, J. P., 2333
*Johnson, C. O., 2584
*Johnson, E., 2514, 2529
Johnson, P. J., 2332
*Johnson, H. L., 2559
*Johnson, W., 2550
Johnston, C., 2175
*Johnston, C. W. 2337
*Jonas G., 2337
Jonas, S. P., 2519 2520
Jones, D. B., 2119
Jones, E., 2314
Jones, F. L., 2371
Jones, H. A., 2510
*Jones, H. I., 2339
*Jones O. V., 2191
Jones, T. D., 2569
*Jones, W. A., 2337
de Jong, H. G. B., 2349
de Jongh, S. E., 2469
*Jonsson A. E., 2494
Jordan, P. O., 2493
Jordi, A., 2483
Joseph, H., 2591
*Josephson, W. S., 2494
Joshi, S. S., 2316
*Jourdan, P. 2330
Joye, P., 2367
Joyet - Lavergne, P., 2473 2174
Jukes, T. H., 2464.

- Loewinson Lesaing, P. 2392
 Loomis N P 2558
 Louler I 2497
 Lowry 1 2427
 Lowkot, K 2541
 Lowndes I C 2512
 Luboscher B R 2378
 Lucas I 2477
 Lucchini P 2402
 Luckash M 2473
 Luke M 2159
 Luth P 2495
 Lutschits H 2538
 Luthmann P 2564
 Lukács J 2488
 Lukens C 2483
 *Lumsden D 2578
 Lumsden T 2481
 Lundsgaard P 2488
 Lunge I 2390
 Lusk G 2462
 Lustig H 2179
 Lutovionov I 2551
 Lutwak C 2471
 Lyaskan B A 2384
 *Lynn A H 2513
 Lynch P W 2474
 *Lyndon R 2371
 Maag W 2443
 McAdam D J Jr 2401
 McCay, C M 2199
 McClelland P D 2351
 McCloud, J L 2399
 McConlham C H 2457
 McCollum, F V 2464 2466, 2481
 *McCullin A A 2551, 2556
 McLarlane, W D 2461
 McGrail P J 2102
 Machado P 2319
 Machatschek P 2390
 Mefferty P W 2450
 McKay, P S 2476
 McKay, J W 2591
 McKay, R W 2318
 *McKeefe P P 2'67, 2568
 *Mackenzie R W R 2578
 Macklin J C 2501
 McKinney, D S 2499
 McKnight, C 2101
 *McLellan A 2590
 McLennan J C 2140 2311
 McVally W F 2103
 McQuinn H W 2379
 McVetty, P G 2379
 Macy, I G 2165
 Madlers, K 2474
 *Madorsky, S L 2407
 *Madruck, Ges für Maschinelle Druck-entwerfung m. b H 2419
 Magassy G 2481
 *Magnesitipar Ltd 2537
 Magnus, P 2509
 Mahadevan C 2359
 Mahant P C 2310
 Mahata V 2525
 Maibe A 2412
 *Main R & A, Ltd 2339
 *Makarov, V L 2350
 *Maklenkov B R V 2541
 Malherbe I de V 2505
 Malinberg T 2393
 *Malvarezskii V I 2530
 Malvarov K L 2196 2387
 Maluca S 2190
 Mandat B K 2499
 Mangold J I 2496
 *Manley R P 2537
 *Manlove Abbott & Co Ltd 2333
 Mann T 2473
 Manning P D V 2521
 Mansons Bak J I 2490
 Mantell C I 2368
 Marcus J K 2468
 *Mansel J W 2536
 Maerens I D 2452
 *Marinetti R 2338
 Marine D 2183
 Mark H 2310
 *Markshafsk B 2551
 *Martini A R 2503
 Martin G 2544
 *Martini P 2476
 Martini V 2431
 Marton I 2481
 Marwick, T C 2312 2319
 *Mars P 2408
 Marx T 2588
 *Maschinenfabrik Oerlikon 2501
 Masaki O 2362
 Mascherpa P 2185
 *Maschinenfabrik Admif Viebahn 2107
 *Maschinenfabrik Augsburg - Normberg, A-G 2338
 *Maschinenfabrik Luckau R Woll A-G 2335
 *Masou, N 2336
 Mathews N L 2565
 Mathieu 2113
 Mathieu (Mme) 2313
 *Mathison, L 2536
 Matignon C 2341
 Matson, P M 2553
 Matsumoto H 2511
 Matsumoto, N 2511
 Matsumura Z 2416
 Matthes, K 2450
 Matthews P J 2542
 *Matuschek J 2374
 Maundley C 2461
 Maume, I 2542
 Maurer K 2498
 *Maurer, K 2494
 May, P 2538
 *Mayer, M 2549
 Mayrann H S 2461
 *Maynard C E 2596
 Mayrhofer, H 2443
 Maurer J 2313
 *Mead Pulp & Paper Co 2564
 Mecke, P 2539
 Mecke R 2563, 2564
 Meersseman P 2485
 Meisera W I 2360 2363
 Megrad R 2474
 Meier F W 2386
 Meljeringh W 2356
 Maerens, K L 2404
 Mellanby, J 2448
 Melton J 2390
 *Melville P 2335
 Mendel, H 2589 2590
 *Mensdorfer K 2377
 *Menschick, W 2471
 *Mentzel A 2550
 *Merck P (Luzma) 2521
 *Metallges A G 2377 2411 2528 2529 2581
 Meunier K 2503
 Meurs G J van 2475
 Meyer J 2476
 *Meyer P 2537
 Meyer Busch R 2485
 *Michel P 2491
 *Michel E 2313
 *Michel R 2359
 *Mihlerrmann H 2408
 *Mihlavacca A 2470
 *Miser P 2576
 *Miska J 2385
 *Mittas R W 2411
 *Mittas H R 2475
 *Mittas A P 2498
 *Mittas P 2335 2380
 *Mittas G 2335 2380
 *Mittas W B 2415
 *Mittas A G 2372
 *Mittas M G 2507
 *Miklavskii N M 2521
 *Mittas I B 2331 2584
 *Mittas P 2351
 *Mittas H 2410, 2511
 *Mittas C R 2390
 *Mittas C R 2567
 *Mittas S 2472
 *Mittas P C 2478
 *Miyamoto, S 2354
 *Miyoshi, S 2560, 2561
 *Modovan I 2478
 *Modoff, I C 2109
 *Mohlman, P W 2502
 *Moles P 2313
 *Moller I 2166
 *Moller I 2159
 *Monasterio, G 2487
 *Monahan, W 2331
 *Monkhouse, A C 2547
 *Monroe C P 2462
 *Monroe R R 2347
 *Moore H H 2462
 *Moppett W 2360 2419
 *Morelle J 2163
 *Moretti P 2487
 *Morgan, A I 2465
 *Morgan G T 2542, 2550
 *Morgan, J D 2546
 *Morgan, J S 2510
 *Morgan W W 2360
 *Morrin N 2479
 *Morrin S 2479
 *Morrow, J A 2513
 *Morrill, G 2364
 *Morton, H T 2103
 *Morton W A 2536
 *Moser, J R 2336
 *Moskowitz M 2517
 *MOTOR Fuel Proprietary, Ltd 2550
 *Motte, M H 2509
 *Mongey, H C 2554
 *Mouriquand G 2467
 *Mouharro, I V 2409
 *Mocharov, V A 2409
 *Munslow, W 2471
 *Mudge, W A 2403
 *Mühlbauer M 2143
 *Müller, C H P A G 2336
 *Müller P 2371
 *Münz P 2440
 *Mukarje J B 2359
 *Mukhyama, M 2368
 *Multicolor Films Inc 2318
 *Mumford W S 2567
 *Munchka J 2501
 *Murakawa K 2361
 *Murnaghan H 2561
 *Murek K 2379
 *Murphy A R 2575
 *Murphy I A 2506
 *Murray Rust D M 2151
 *Murray D R P 2418
 *Murray I V 2391
 *Musolino G 2487
 *Myers J T 2559
 *N. V. de Botsaafche Instuut voor Maas-achappij 2478 2480
 *N. V. Maatschappij tot Baas en Exploitatie van Oer-ruwen 2536 2537
 *Naegeli C 2471 2472
 *Naerli H 2374
 *Naffel J A 2509
 *Nagy I 2520
 *Nakahara, W 2167
 *Nakamura S 2474
 *Nakai T 2446
 *Nakata, A 2351
 *Nakhmanovich M I 2585
 *Nanji D R 2377
 *Narasimamurti, N 2456
 *Nattayan T S 2316
 *Navini, A G 2356
 *Navini, R 2393
 *Neethaus, H 2368
 *National Processes Ltd 2106
 *Natta G 2342
 *Naves, L 2412
 *Naylor, W R 2108
 *Nealay, J R 2527
 *Needham J 2452
 *Neulich S A 2532
 *Neill J M 2482
 *Nelson, J 2567
 *Nelsons Salk, Ltd 2567
 *Némec, A 2507, 2508
 *Némova Z 2392
 *Nernst W 2556
 *Nervi C 2481
 *Neuberger, L 2507
 *Neue Element Werke Gebrüder Haas & Co 2374
 *Newman, R K 2440
 *Nichols, R H 2493
 *Nichols M S 2502, 2504
 *Nicolaou, I 2478
 *Nielsen, H 2550
 *Nielsen R 2303
 *Niess H 2405
 *Niethammer, A 2388, 2459
 *Nikulin, B A 2359
 *Nilsson R 2447
 *Nims B 2465
 *Nishida K 2419 2509

- Vichimura S 2449
 Vittinger C 2540
 Vitzum I 2444
 Viven A 2496
 Voerdinger H 2518
 Voper 2446
 Nolan P J 2550
 Volze H 2494
 Vonnara E 2449
 Voordt Ltd 2497
 Vorris G C 2497
 Vorris W 2593
 Vorthrop J H 2450
 Vorse T 2474
 Vuckols A H 2555
 Vugent W W 2549
 Vauwener M 2444
 Vettial W H 2593
 O'Brien T E H 2593
 Obet W 2512
 Odda Smethverkt Akt
 selskabet 2514
 2529
 Odell W W 2550
 Ohman E 2400
 Omer H 2472
 Oesch J B 2575
 *Osterreichische Chem.
 Werke G m b H 2550
 *Osterreichische Sacm-
 ens - Schachtel -
 Werke 2538
 *Ofeldt, P W 2537
 O'Flaherty P 2548
 O'Jacks S 2459
 Oboto H 2532 2458
 2549
 Oka S 2547
 Okagawa Y 2448
 2449
 O'Leary M J 2542
 2545
 *Ola, Princess sur
 Lippe 2477
 Oliver C V 2564
 Oppel A R 2578
 Osood A R 2543
 Osmi L 2509, 2510,
 2511
 *Oswanburger Chemi-
 sche Fabrik A-G 2429
 *Ordody, B 2564
 *Ordody, L 2566
 *Ordelap J W 2533
 Orent E 2464
 *Ormesher P L 2540
 Ormott B 2556
 Osann B 2405
 Osgood E E 2453
 *Osmuthu V P 2550
 *Oster B 2536
 *Osterday G A 2541
 Osterhout W J V 2541
 *Osthoff W 2577
 Ostwald W 2544
 Otam S 2443 2445
 O'Toole E 2444
 Ott E 2551
 *Otto C & Co G m
 b H 2541
 Otto K 2477
 Overhoff J 2462
 *Owe A W 2524
 *Owens A G 2574
 Owens J S 2553
 Owens E 2525
 Pace P T 2446
 *Pace A 2410 2411
 *Paffen W 2494
 Page I H 2471,
 2485
 Paid 2543
 Palant A I 2521
 Palmer F Jr 2537
 Palmer L A 2494
 Palomaa H H 2417
 Pan L C 2570
 *Pan American Pe-
 troleum Co 2557
 *Panhandle Refining
 Co 2544
 Pansin P 2554
 *Paper Patents Co.
 2467
 *Pepkov V V 2530
 Pepp F 2591
 *Parent A C J 2525
 *Parfenov V V 2555
 *Park C R 2550
 Parmelee C W 2535
 Parravano N 2535
 *Paruch J 2517
 Parsons P B 2492
 Partington J R 2532
 *Paschke M 2575
 Pascoe E H 2524
 Paskaj J 2512
 Passano R F 2494
 Passerini M 2423
 Pasternak L 2445
 *Patent - Technische
 Ges für Elektrische
 Glühlampen 2534
 Patwardhan V N 2450
 *Paul R 2493
 *Pavlovich P L 2550
 Pavolov L 2554
 Paya M 2543
 Pearce J W 2544
 Pearson T C 2532
 Pécet A 2571
 Peppers E 2594
 Pelzer J 2539
 Pepin - Lachance, J 2537
 Péru M 2494
 *Perry E W 2550
 Perreboom S S 2507
 *Perthois E 2442
 Perkins A E 2492
 Perman E F 2549
 *Permut A G 2515
 Perrot E 2523
 *Peters J M 2547
 Petch C E 2512
 Peter A 2545
 *Peters K 2597
 Peters R A 2445
 *Petersen L S 2540
 Peterson W H 2445
 Petri L 2513
 Petro V A 2465
 *Petrov A K 2531
 *Petrov G S 2531
 2544
 Petryanov I 2540
 Pettit A E J 2542
 Petty E 2543
 Petryjohn E S 2544
 Pfabi C 2544
 *Pfeiler L 2497
 Pfeiffer G 2471
 *Pfeiffer J 2579
 Phelps P B 2503
 Phelps S M 2534
 Phrasom, T 2444
 *Pick P 2535
 *Pickard J A 2537
 Pickard R H 2549
 Piersanti J 2543
 Pigott M G 2463
 Pilar F 2474
 Pilon G 2454
 *Pilot Laboratories
 Inc 2493
 *Pine L A 2514
 *Pine F 2577
 Pines L 2453
 Pinto M D 2591
 *Pintch, J A-G 2550
 Pisk F 2559
 *Pisnev Z V 2504
 Planchon R 2504
 *Platen - Munster
 Refrigerating Sys-
 tem Aktiengesell-
 schaft 2535
 *Pérea L 2497
 Pocking J 2448
 Pohl E 2579
 *Pohlmann J 2493
 *Pohlmann V 2579
 Pold G R 2535
 *Polodnik F 2553
 *Poland R C 2499
 *Pope W J 2524
 *Pora - Werk Paul
 Schrot 2550
 Porter A W 2554
 Potthmann W 2499
 Potts T 2544
 Powers T F 2500
 Pratt H E 2545
 Preiss M J 2554
 *Preussische L 2554
 Preston F W 2534
 Preston G D 2597
 Preston L M 2445
 *Preuss K H 2593
 Price W A 2552
 *Priest R 2579
 *Priest J 2532
 *Priestley E 2431
 *Pritchard W Jr 2569
 v Probstner A 2472
 Prost E 2592
 *Proyanov, S P 2559
 *Pross M 2502
 *Proyekt 2548
 *Pucher G W 2454
 Puchols P 2494
 *Putscher J J 2491
 *Quarts et Söde 2537
 R N 2524
 Raab F 2499
 Rabbano A 2454
 *Rabmann H 2535
 Radnowicki J M 2474
 *Rademaker Ma-
 schinenfabrik August
 Koenig G m b H 2538
 *Radomsky Parent
 Ltd 2574
 Radjhan T C 2518
 Rae G 2591
 *Rauch C 2532
 Raitt W 2544
 *Raymond F 2574
 *Ramage A S 2541,
 2542 2557
 Ramart - Lucas
 (Mime) 2567 2415
 Rammer E 2545
 Randall M 2593
 *Randolph D W 2572
 Ransom A 2494
 Ransoff G 2467
 Rapatz P 2494
 Raper H B 2470
 Rastu, P 2565
 *Rawers, J R F 2493
 Rawling S O 2574
 Ray B H 2560
 Ray R C 2547
 Razova P 2535
 *Razumeyev, A A 2575
 Re U 2545
 Reckendorfer F 2537
 Reclaire A 2521
 *Red River Refining
 Co, Inc 2559
 Reed G H 2544
 Reed L J 2534
 Rees O T 2590
 Regnier E 2539
 Reihman W C 2534
 Reichard S K 2555
 Reid E 2531
 *Reid M 2594
 Reiner M 2516
 *Reiser O 2495
 Reinhardt L 2549
 Reinhold H 2543
 Reisinger H 2539
 *Reinisch E 2517
 *Reinmann W 2540
 *Remington A M 2499
 Remington R E 2455
 Remondie 2412
 *Reish M F 2540
 *Respa Inc 2568
 *Reuss W 2411
 Reuter P 2479
 Reuther A 2500
 *Reut C R 2529
 Reysche J 2505
 Ricard R 2542
 Riehl A 2479, 2480
 Riehl P 2479
 Riehl T 2535
 Rice G 2594
 *Richardson H T 2495
 Richardson S 2564
 Richarz H 2539
 *Richmond A 2558
 Richter F A 2591
 *Richter G, Ltd 2442 2524
 Richter L A 2593
 *Ricks S J 2504
 *Ried O 2567
 Riedig P 2527
 Riepe J 2537
 Rijn, J J L van 2461
 Rinek E 2535
 Ringel A 2504
 Rips R 2492
 *Ripley P C 2411
 Ripperton J C 2587
 Riss G 2566
 Ritter P 2526
 Ritter H 2545
 *Ritterfeld W 2532
 Rives L 2454
 *Road Development
 Co Ltd 2538
 Robb J A 2544
 *Robbins T 2497
 Roberts J 2542
 *Roberts R P 2567
 *Robertson R N 2535
 *Robinson E B 2597
 *Robinson E H 2578
 *Robinson H N 2539

- *Robinson W. J. 2337
2350.
Robinson P. I. 2382
Robinson, H. L. 2492
Robinson, W. C. 2504
Roddie, W. T. 2359
Rode, E. V. 2383
*Rodionov A. M.
2350
Roe J. H. 2491
Roethel H. H. 2404
Koffin A. H. 2478
Rogers R. 2476
Rogers T. H. 2354
Rosenfeld C. 2334
Roholm A. 2443
*Roi-Letter et Cie
2340
Rolet, A. 2493
Kull, L. J. 2417
Ruller, P. R. 2343
Rim, P. 2318
Roman F. L. 2504
Romig J. W. 2347
*Romualte A. 2348
Romeo A. 2469
Kok, M. 2402
*Rose J. 2341
Kosenbom, A. 2447
*Rosenqvist, C. 2373
Roussal, N. 2340
Rous, P. 2343
Rosa, H. 2477
Rosa, J. H. 2362
Rosenbach, M. L. 2339
van Rossum A. 2373
Rond, A. 2421
Rond L. 2478
Rosenman J. 2364
2360 2393
von Roustecy, E.
2475.
*Rouveau, L. F. J.
2374
Routa's C. 2335,
2383
Roux A. 2403
Rovesti G. 2422
Rove, P. M. 2424
Rowland, H. R. 2373
*Rowlands T. 2400
Rowntree R. K. 2402
Knee, H. D. 2370
*van Royen, H. J.
2401
Rosenkrantz, L. 2338
*Rubber Service Lab-
oratories Inc. 2394
*Rodolfo, W. 2504
Rueff H. 2339
*Ruff, O. 2376
Rupp R. 2338
Russell A. C. 2332
Russell A. R. 2483
Ruth, H. D. 2479
Rutler, J. F. Jr.
2488
*Ryane A. 2379

Sach W. 2338
Sadtler, S. D. 2390
Sagbarth B. 2333
Sahlin T. G. 2391
*A. I. G. (See anon
Inventories Catalog
nin) 2498
Saito, M. 2450 2491
Sakata, H. 2446
Sakima B. 2352
Sahlert, F. 2314
*Salerno M. 2498
Calleras J. 2483
Salman J. 2336
*Salmon Karsten A.
2504
Salomon, H. 2489

*Samaruthee R.
2337
*Samuram D. 2333
*Sample L. 2337
Sanches J. A. 2399
*Sanford H. 2333
Sanford H. N. 2444
Sann M. 2452
Sano T. 2341
*Santana J. 2350
Santana A. 2411
*Sany M. 2406
Sasaki K. 2371 2397
Sasaki R. 2494
Saschek W. J. 2343
Sauter P. 2337
Sausy L. 2339
*Savitsky, M.
2387
Saxon, R. 2371
Searpa O. 2368
*Searwith G. D. 2308
Schaal G. 2448
Schaack H. C. 2336
Schäfer C. 2433
*Schaffer L. 2379
Schaffer J. 2444
Schauer W. 2443
Schaur R. 2408
2409
Schaffel 2331
Schaffers H. W. 2333
Schäfer J. 2360
Schamel 2494
Schreck, H. 2394
Scherrer I. 2339
Schering Kahlbaum
A. G. 2437 2379
2447, 2413 2421
Schering G. (Girma)
2340
Schick F. 2336
Schicke W. 2432
Schickels T. 2331
Schickold, 2339
Schiedek O. 2336
Schiffner, H. J. 2407
Schikner, G. 2343
Schiller A. 2336
Schlingens R. 2332
Schinke P. 2408
Schirmacher, K.
2441
Schittenhelm A. 2444
Schloss, H. M. 2474
Schloss, R. H. 2340
Schmidt R. 2362
Schmidt D. 2374
Schmidt, H. 2374
Schmidt, J. 2383
Schmidt R. 2323
Schneider, A. 2494
Schneider, F. 2353
Schneider, H. 2377
Schneider, K. 2332
Schnitzspahn, L.
2429
Schuber, H. 2361
*Schubert, C. J.
2337
*Schun, I. 2323
Schunberg, A. 2443
Schunberg, A. 2336
Schöfeld H. 2362
Schöfeld R. K. 2339
Scholts, H. G. 2471
Scholz, H. 2399
*Scholz, V. 2378
Schottak, A. 2338
*Schotte, H. 2439
Schraab, K. 2324
Schrimmacker, F.
A. H. 2446
Schrempf A. 2348
Schroder, C. 2341

*Schroeter G. 2441
Schubert C. I. 2371
Schulek E. 2319
Schultze G. R. 2373
*Schumacher K.
2406 2409
Schumann, T. R. W.
2342
Schussch H. L. 2335
Schuster H. 2344
2343
Schuster L. 2339
Schwall C. L. 2333
*Schwarz H. 2331
Schwarz R. 2379
Schweert K. 2448
Scott R. H. 2334
Scott T. J. 2436
*Scottish Dyestuffs Ltd.,
2438 2374 2373
Scott G. 2464
Seeman L. 2466
2467
Seeman Co. The
2363
Seach W. T. 2338
Seales, I. 2400
Seisell L. H. 2308
Seisell W. J. 2337
Seitz F. 2337
Seely H. H. 2339
Selito N. 2337
Selz C. G. 2370
Seligman R. 2494
Sendinger optische
Glasswerke G. m.
b. H. 2336
*Seufert G. 2494
Seeger J. F. M. 2336
Seerick N. 2473
Seeth J. R. 2361
Seether, L. R. 2362
Seesay, A. 2478
*Seifer, J. J. 2329
Shakhnoork, I. G.
2336
Sharp, V. A. L. 2488
*Shastri H. J. 2333
Sharrow N. P. 2390
Shaw M. H. 2362
2363
Shar, H. 2474
Shchekokhin, N. G.
2390
Shel R. W. 2352
Shells, O. 2483
Shelton, S. H. 2409
Shelling D. H. 2464
Sherman, H. N. 2468
Sherwood, F. W. 2482
Sherwood T. K. 2493
Shichiri, C. 2443
Shields N. J. 2350
Shimohara H. 2311
Shimoda, Y. 2362
Shirabe Y. 2446
*Shivers P. F. 2336
Shirner R. L. 2417
Shirner, L. C. 2431
Shuklin, N. I. 2420
Shukler, C. P. 2463
*Shulden, H. L.
2304
*Shurigin, P. N.
2390
Shalya L. 2363
Shigwick N. V. 2345
*Shimens F. A. H.
2336
*Shimens, F. C. 2408
*Shimens & Haldre
A. G. 2337, 2373
2376, 2408 2310
*Shimens-Flaschenwerke
A. G. für Kohlen-
fabrikate 2376.

*Shimens Reiniger
Verfa. Ges. für Mech.
anische Technik
2335
*Shimens Schoket
werke A. G. 2374
2377 2409 2377
Shims, A. B. 2309
*Shirp P. 2304
Shivers A. 2314
Shivron A. 2341
Shitonen T. A. 2417
Shitonen T. A. 2494
*Shirman A. 2377
*Shirshov A. 2363
Shimane A. 2343
Shimon J. H. 2341
Shimon J. L. 2433
Shimon J. L. 2334
Shin P. C. 2367
*Shir I. R. See Itali-
ana Kierische Indus-
trial 2350
*Shull V. 2448
Shuller M. H. 2401
Shuler J. 2376
Shimono N. 2341
Shlander C. J. 2370
de Smidt J. 2362
Smirnov D. I. 2334
Smith A. 2357
Smith, A. E. 2421
Smith C. R. 2401
Smith F. C. 2476
Smith F. C. 2374
Smith I. W. 2443
Smith L. W. 2442
Smith M. P. 2463
Smith, R. L. 2464
Smith S. H. 2353
Smith W. 2456
Smith, W. 2434
Smith, A. 2353
Smorodintsev I. A.
2441
*Snyder R. W. 2397
Solovik, H. 2316
*See anon des ac-
cumulateurs mono-
plaque, 2374
*See anon Akulina
2329
*See anon des ancien-
s'installations
Skoda A. H. 2410
*See anon des ap-
pareils de manuten-
tion et fours stein
2336
*See anon des distil-
leries des Deux
Sevres, 2317, 2439
*See anon des hauts
fourneaux et fonderies
de Pont-a-Mousson
2408
*See anon, pour l'ind-
ustrie à Râle 2331,
2372, 2374 2373
*See anon M. Naef &
Cie, 2437
*See chimique de la
Grande Paroisse
Arore et produits
chimiques, 2497
*See française des pro-
duits alimentaires
Arores 2497
*See Jeanjean et
Cassas, 2364
*See de recherches et
de perfectionne-
ments industriels
2341.

- Soc des unnes chim
 igues Rhône
 Volens 2437
 Societä altformand
 erie acierierne e
 ferriere Franchi
 Gregoris 2407
 2408
 S. Hoor K 2341
 Sokolova N 2344
 Solariu G 2466
 Sommer A 2468
 Sora F 2474
 South Metropolitan
 Gas Co 2336
 de Souza Vieira O
 2340 2333
 Spackman I S 2460
 Späker K 2708
 Speiser A K 2463
 Speitel R 2426
 Spence F & Sons,
 Ltd 2332
 Spicers Ltd 2374
 Spiller P 2334
 Spindel M 2334
 Spitzer K 2331
 Spitznagel V I 2341
 Spoelstra D H 2331
 Spoon I W 2452
 Sprenger (I) Patent
 verwertung Jrothka
 m B H 2404
 Spring L W 2409
 Spruijer F 2334
 Spruijt J P 2407
 Sreenivasaya M,
 2456
 Stacy H R 2354
 Stadlinger H, 2371
 2390
 Stednicko T 2353
 Stiege A 2441
 Stähler F 2444
 Stähle R 2393
 Stahl W 2334 2336
 Stalder W 2352
 Stamatenko B 2404
 Stamberger P 2302
 Stanbery L J 2402
 Standart H C, 2427
 Standartwerke Ros
 tock & Baerlöcher
 2497
 Stanley W E 2402
 Stark T J 2370
 Starr J B 2495
 Stary Z 2473
 Stasak A 2345
 Staton J C 2494
 Stauffer H 2491,
 2492
 Standt F 2493
 Stechow N 2461
 Steen K E 2464
 Steenbeck M 2357
 Steenbeck H 2464
 Steger A 2383
 Steierwald F 2446
 Stein F 2410
 Steiner R 2448
 Steiner W 2362
 Stephenson H F
 2348
 Sterling V 2366
 Sterne E T 2339
 Stevens R H 2364
 Stevens R H 2407
 Stevens W O 2359
 Stewart F C 2460
 Stewart O M 2357
 Stosny E 2390
 Stoppel C 2349
 Stoppel C 2439
 Stalwell G R 2378
 Stummel R M 2499
 Streckhager D C
 2360
 Stockfisch H L 2371
 Stockton A B 2464
 Stokas A 2404
 Stoklasa J 2456
 Stolpp T 2413
 Stommel 2301
 Stover K C 2424
 Stots R 2394
 Stoughton H 2401
 Stout L E 2370
 Strachan I 2465
 Strachan M T 2440
 Stranski I N 2341
 Strauß F 2454
 Strauß W 2454
 Strobl I J 2374
 Struve H 2360
 Stricker G B, Jr
 2494
 St. Subner F 2466
 St. 1 2377
 Sturrock J L 2351
 Succop J A 2364
 Sumner J 2473
 Susslen G F 2337
 Suez J A 2442
 Sulzmann R 2464
 Sula H 2440
 Sullivan M X, 2449
 2470
 Sumi M 2467
 Sumiki Y, 2457
 Sunderman F W
 2473
 Sundström F F
 2464
 Superheater Co Ltd
 2449
 Sure R 2464
 Sutton P 2390
 Suzumura A 2350
 Szaboda J 2402
 Swarts F 2339
 Swarts J H, 2462
 Sweetland E J,
 2335
 Swends N, 2497
 Swiss Inventions An
 dicate Ltd 2340
 Swope H G 2303
 Szabo T 2325
 Szählender K 2333
 2344
 Szakacs O 2491
 Szekeres G 2444
 Szent György A 2484
 Székely J F 2457
 Székely T 2384
 Taitner M L 2464
 Taisak K A 2415
 Takagi T 2448
 Takahashi I 2411
 2412
 Takahashi T 419
 2440
 Talbot J H 2469
 Talwalker T W
 2335
 Tanaka R 2346
 Tandler R 2490
 Tanemura A 2360
 2461
 Tankó B 2461
 Tarabov K A 2507
 Tarasov K I 2431
 Tari O F 2428
 Tarugi N 2452
 Tassik K 2411
 Tate W R 2444
 Tatsu M 2446
 Tatu H 2371
 Teubenhau M 2470
 Teuber H 2444
 Taylor K A 2361
 Taylor Wharton Iron
 & Steel Co 2410
 Techner J 2444
 Tedham W F 2334
 Temesváry, I 2420
 Terenin A 2354
 Teen R, 2454
 Torres P 2334
 "Torvalper L O
 2407
 "Tortiles Artificio
 lve A Maurer
 2467
 Tervinsky F 2339
 Thamm I 2352 2469
 Thallie V N 2363
 Thaa A 2464
 Thero E R 2466
 Thernault, J, 2354
 "Thermotectic Ltd,
 2335
 Thomas A A, 2339
 "Thomas, J 2434
 2374
 "Thompson A P,
 2331
 "Thompson R F 2374
 Thoreau J 2442
 "Thornton J H, 2374
 Thornton S F, 2403
 Thummet W G 2370
 "Thurner F 2464
 Thurner C F 2339
 "Thuring O O 2350
 "Tichy G 2407
 Tilfer F W 2464
 Tisdale G B 2492
 Todd C, 2416
 Todd G B, 2355
 Tomkinson M 2445
 Tool A O, 2354
 Torrey, H J 2354
 "Torrence H 2354
 "Townsend L W R
 2355
 "de Trarup M K,
 2375
 "Tramer L 2441
 Trechsch R, 2370
 "Trebits G 2337
 Triang O 2344
 Triault G 2443
 Trier G 2423
 Trifonov F I 2404
 Trivich A. P H
 2378
 Trochin W 2342
 Trotman E R 2466
 Trotman S R 2466
 "Tschew V 2497
 Tscholakov J 2413
 Tsunomura I 2445
 "Tutor Accumulator
 Ltd 2338
 Tu F C 2446
 Tukats S 2354
 Tuomann O 2461
 Tupholme C H S
 2334
 Turner C W 2474
 Turner D 2336
 Tuason P 2433
 Tvers D F 2345
 "Twa D F, 2496
 Ubbelohde L 2352
 Uchevsk Opitnoi
 Zvod Moskovskago
 Tekhnika Zhuro
 vPromyshlennosti
 "Tekhnikaavud,"
 2381
 Ueno S 2352
 "Ude F 2357
 "Ukrainsk Nauchno
 Isledovatel'skaja
 Institut Metallurg
 2378
 "Uman G 2357
 "Ulrich A L 2331
 "Ulrich F 2349
 Unger W 2366
 "Union chimique belge
 2315
 "Union chimique belge
 S A, 2496
 "Upton R G, 2402
 "Urbain E 2314
 "Ure W, 2355
 "Urbant, J W, 2397
 "Urbant, F W, 2413
 Vadász R 2391
 Vallette 2370
 Valica V, 2353
 Valance R H 2353
 Varroen E 2349
 Vas, K 2491
 "Vaskelyi, J 2452
 "Vale R 2467
 "Vasni I L, 2375
 "Vasny L, 2317
 "Vascha A B, 2360
 "Vasch A H 2397
 "Veldre M A 2477
 "Vendi A, 2392
 "Venkateshwar, B,
 2343
 "Verdine, A, 2333
 "Verona für chemische
 Industrie A. G.,
 2467
 "Verriete Aluminu
 m Werke A. G.,
 2377
 "Verriete G f h
 Lampen und Elec
 tricitats A. G. 2376
 "Verriete Stahl
 werke A. G. 2407
 "Verriete Telephon
 und Telegraphen
 fabriks A G Cerny
 Noid & Co 2410
 "Verkaufsvereinigung
 für Feinverpackung
 G m B H 2354
 "Vermorel V 2314
 "Vernadsky V I 2359
 "Verrill, J 2352
 "Viale G 2453
 "Vaslet J 2377
 "Vasari H 2455
 "Vickers H B, 2458
 "Vierheller H 2338
 "Vinickaya R V,
 2375
 "Vinkov M A,
 2404
 "Vitáls I 2444
 "Vitez D 2340
 "Vladimirov, K. V
 2410
 "Vlassopoulos, V 2463
 "Vlassov, S 2381
 "Völcker E 2336
 "Volkovich S, 2326
 "Volkovich S I, 2407,
 2304
 "Volkov N S, 2366
 "Vondráček O, 2343
 "Vondráček, J, 2329
 "Vau Voichis, C. C.,
 2371
 "Vorobkov, N V,
 2376
 "Vorobitsov N N,
 2437
 "Voss W, 2329
 "Vostokov A I, 2345

- *Vostokov, N. A., 2335
 de Vries N. P. 2310
 *Vuk M. 2193
 *Vulcan, Feuerung
 A. I., 2329
 Wala K. 2324
 *Waldell M. 2377
 *Wale G. & Son
 Ltd 2337
 *Wale G. A. 2337
 *Wale H. 2370
 Wassman S. 2393
 Wagner A. 2394
 *Wagner H. 2337
 Wagner I. R. 237
 Wagner H. 2153
 Wagner J. Ph. 2399
 Wagner Jautz T.
 2418
 *Wahl A. S. 2403
 *Walach H. 2142
 *Waller K. 2341
 Wallachmuller Lutz
 P., 2116
 Walker J. A. 2361
 Walker O. 2456
 Walker T. K. 2316
 Wall P. J. 2374
 Walls P. J. 2397
 Walls L. P., 2370
 Walstead, J. P. 2397
 *Walter, G. 2197
 Walther, A. K. 2312
 Walther K. 2164
 Wamischer, L., 2161
 Wangenstein, O. H.
 2174
 Wan, H. P., 2354
 Warburg O., 2431
 2172
 Warran, D. M., 2194
 Warren, W. H., 2114
 Warth, F. J., 2161
 Wasmuth R., 2161
 Wawermann A., 2143
 *Watanabe H., 2323
 *Watanabe T. 2154
 Waterhill R. W. J. 203
 *Waterhouse, W. P.,
 2323
 Waterman H. I., 2370
 Watson H. A. 2474
 *Watson, H. C., 2377
 *Watson, W. N., 2107
 Watts, O. P. 2370
 Weber F., 2474
 Weber, H. 2700
 Weber, B., 2443
 Webster J. P. 2154
 Welekind P. 2132
 *Weger A., 2337
 Welchert C. K. 2473
 *Weigel A. 2350
 *Weil K. 2367
 Weinreb P. 2333
 Weinstock M. 2165
 Weinstock Z. 2321
 Weirick H. C. 2347
 Weiss 2401
 *Weiss J. M. 2381
 *Weinrich A. 2374
 Welch H. 2474
 Weld L. R. D. 2357
 Welker W. H. 2474
 Werken 2324
 Wertheimer P. 2143
 Wesson I. C., 2161
 West G. H. 2393
 West W. 2464
 *Westbrook J. R.
 2375
 *Westliche Mas
 chinenfab. Ges.
 C. m. b. H. 2351
 Weyer F. 2356
 Weyl H. 2367
 Wheeler C. P. 2502
 Wheeler J. A. 2363
 Wheeler M. A.
 2346
 Wheeler R. V. 2353
 *Wheeler T. L. 2331
 Whipple R. S. 2379
 *Whitby G. S. 2334
 Whisby I. 2101
 White H. L. 2112
 White P. 2347
 *Whitehead L. D.
 2107
 *Whiteman G. H.
 2111
 Whiton, H. S., 2501
 *Whittemore Hul
 bert Whittemore
 A. Belknap, 2374
 *Wickwire Spencer
 Steel Co., 2410
 *Wiegand F. L. 2411
 Wiegand, G., 2316
 2317
 Wiener P., 2500
 Wiert, R., 2354
 Witch, C. C. 2354
 Wilcox, A. R., 2520
 *Wilhelm, J. O., 2310
 *Wilhelm, R., 2351
 *Wilkinson, J. A., 2342
 *Williams D. A., 2332
 *Williamott S. G.,
 2134
 Willner P. 2113
 *Willshaw H. 2197
 *Willstätter R. 2194
 *Wilms L. O. 2313
 *Wilson P. G. 2390
 *Wilson H. J. 2334
 *Wilson J. A. 2547
 *Wilson R. P. 2156
 *Wilson S. D. 2313
 *Wilson W. C. 2341
 *Winisch G. 2497
 Wintermiz R. 2475
 Winterstein A. 2424
 Winterstein H. 2423
 Wirsch A. M. 2474
 Wiseman B. J. 2144
 Witte P. 2370
 *Wolner A. 2497
 Wolf I. G. 2370
 *Wolf K. 2374
 *Wolfs R. A. 2372
 *Wolff P. 2311
 *Wolf I. K. 2163
 *Wolfram R. 2331
 *Wolfram M. 2370
 *Wollenberg H. 2334
 *Woodbridge J. L.
 2374
 *Woodcock J. W. 2311
 *Woodlands Ltd
 2193
 *Woodman A. G., 2197
 *Woodroffe D., 2344,
 2349
 *Woodson J. C. 2344
 *Woodward I. A.,
 2344
 *Woolley J. G. 2177
 *Wooters W. J. 2343
 *Wort H. K., 2370
 *Worthington, P. V.,
 2363
 *Wortley, G. W. 2339
 *Wosiek O., 2475
 *Wright, C. P., 2351
 *Wright R. I. 2321
 van Wollen - Scholten
 W., 2379
 *Wünsch, G., 2333
 Yamada, A., 2509
 Yamamoto K., 2300
 Yamamoto, Y., 2506
 Yamazaki R., 2542
 Yamazaki M., 2147
 *Yanovsky P. 2347
 *Yorston F. H. 2362
 Yoshida M., 2361
 Yoshikawa, K. 2411
 Yoshimura, K. 2509
 2544
 Zahn R. V. 2199
 *Zahn & Co. G. m. b.
 H. 2329
 Zakschil V. K. 2504
 Zappi I. V. 2352
 *Zarotschenoff M.
 T. 2494
 *Zavertnik J. 2551
 2556
 Zboray B. 2314
 Zborak H. M. 2371
 Zechmeister I. 2133
 Zehnder L. 2379
 Zeile K. 2444
 Zerstet W. 2417
 Zeligson N. P. 2475
 Zelikman I. P. 2545
 Zelnick N. D. 2129
 Zellner H. 2516
 *Zellstoffabrik Wahl
 hol 2564
 Zener C. 2357
 Zenthofer J. A.
 2335
 Zetia I. G. 2442
 Zernik P., 2314
 Zetzsch P. 2155
 Zhikharovich S. N.
 2334
 Zhurik G. I., 2544
 Ziba S. S., 2164
 Ziller, L., 2167
 *Zimmermann & Jan
 sen G. m. b. H.
 2334
 *Zimmermann, P.,
 2104
 Zipperer, L., 2346
 Zita, N. 2423
 Zolinski, J., 2321
 de Zombory, L., 2346,
 2347
 Zouckermann, R.
 2363
 Zscheile, P. F. Jr.,
 2460
 Zulehaninov, V. F.,
 2534
 Zuercher, G., 2169
 Zwickowsky, E. J.,
 2444
 *Zyslov, K. I., 2390

- [illegible]

CHEMICAL ABSTRACTS

Editor: E. J. Crane, Ohio State University, Columbus, Ohio.
Associate Editors: Elmer Hockatt, Hatan Gama and Leonard T. Capelli
Assistant Editors

St. A. Anderson	D. J. Demarest	P. B. Hawk	L. A. Olney	J. J. Skinner
W. L. Badger	Alden H. Emery	P. H. Howe	A. H. Patterson	W. R. Stearns
J. K. Barton	W. D. Emery	C. H. Koser	T. G. Phillips	C. G. Storm
Edward Barrow	W. P. Paragher	H. A. Lepper	A. N. Richards	H. H. Symmes
S. R. Benedict	A. C. Fieldner	R. K. Marshfield Jr.	Richard Rimbach	P. F. Haderhill
P. C. Blau	C. B. Link	A. R. Shults	Alfred Rogers	H. G. Welle
P. I. Browne	C. N. Frey	Harlan S. Mear	C. A. Houtler	C. J. West
C. H. Curran	H. W. Lillie	Chas. H. Mear	A. H. Salin	H. T. Wherry
J. K. Dale	R. A. Corneer	C. R. Sturges	J. P. Schuler	H. P. Wightman
C. C. Davis	W. T. Hall	W. A. Noyes Jr.	H. Scherbel	J. C. Witt

CONTENTS

Patents	1	12. The Fermentation Industries	2204
Author Index	ii	17. Pharmaceutical Chemistry	2207
1. Apparatus and Plant Equipment	2212	18. Acids, Alkalies, Salts and Sundries	2218
2. General and Physical Chemistry	2202	19. Glass, Clay Products, Refractories and Enamelled Metals	2222
3. Subatomic Phenomena and Radiochemistry	2222	20. Cement and Other Building Materials	2222
4. Electrochemistry	2243	21. Fuels, Gas, Tar and Coke	2232
5. Photography	2222	22. Petroleum, Lubricants, Asphalt and Wood Products	2240
6. Inorganic Chemistry	2222	23. Cellulose and Paper	2222
7. Analytical Chemistry	2222	24. Explosives and Explosions	2222
8. Mineralogical and Geological Chemistry	2222	25. Dyes and Textile Chemistry	2222
9. Metallurgy and Metallography	2222	26. Paints, Varnishes and Resins	2222
10. Organic Chemistry	2241	27. Fats, Fatty Oils, Waxes and Soaps	2222
11. Biological Chemistry	2222	28. Sugar, Starch and Gums	2222
12. Foods	2222	29. Leather and Glue	2222
13. General Industrial Chemistry	2222	30. Rubber and Allied Substances	2222
14. Water, Sewage and Sanitation	2222		
15. Soils, Fertilizers and Agricultural Poisons	2222		

Published by THE AMERICAN CHEMICAL SOCIETY, which also publishes *The Journal of the American Chemical Society* and *Industrial and Engineering Chemistry*

Communications for the editor should be addressed to R. J. Crane, Ohio State University, Columbus, Ohio

Remittances for annual dues (\$15.00) in the American Chemical Society or for subscriptions to this *Journal* (\$7.50 per year), changes of address, orders for subscriptions or back numbers, claims for missing numbers and all matters intended for the *Proceedings* should be sent to the Secretary, Charles L. Parsons, Mills Bldg., Washington, D. C. Foreign postage is extra, being \$1.50 per year added to the dues of members residing outside the United States, its possessions and the countries of the Pan American Union and \$1.50 per year for *Chemical Abstracts* alone, except that only one-half of these amounts is added for journals going to Canada. The countries of the Pan-American Union are Argentina, Bolivia, Brazil, Chile, Colombia, Costa Rica, Cuba, Dominican Republic, Ecuador, Guatemala, Mexico, Nicaragua, Paraguay, Peru, Republic of Honduras, Republic of Panama, Salvador, Spain and Uruguay.

The Council has voted that no claims will be allowed for copies of journals lost in the mails unless such claims are received within sixty days of the date of issue, and no claims will be allowed for issues lost as a result of insufficient notice of change of address (Ten days advance notice required). "Missing from files" cannot be accepted as the reason for honoring a claim—Charles L. Parsons, Secretary.

PATENTS

Abstracts of patents are included in the appropriate sections of the journal according to subjects. The abstracts of Australian and French patents are made by William

Published by the American Chemical Society, Easton, Pa. Secretary's Office, Mills Bldg., Washington, D. C. Entered as second-class matter at the Post Office, Easton, Pa. Acceptance for mailing at special rate of postage provided for in Section 1103 Act of November 3, 1917, authorized July 13, 1918. Mack Printing Company, Easton, Pa.

Glass, those of Austrian patents by F. H. Overell, Belgian patents by A. Papineau Couture, British and United States by Earl T. Ragan, Canadian by James T. Mitchell, Danish by Arne Drogseth, Dutch by B. J. C. van der Hoeven, German by William Glass, H. C. Miller and F. H. Overell, Japanese by K. Kashima, Norwegian and Swedish by C. H. A. Robak, Russian by A. A. Roehthigk and Swiss by H. C. Miller and F. H. Overell.

Copies of any patent can be obtained for United States, 10 cents, sent to the Patent Office, Washington, D. C., Australian, 1s 6d, sent to the Commissioner of Patents, 581 Bourke St. Melbourne. Austrian, 1S, sent to the Österreichisches Patentamt Vienna. British 1s plus postage (usually 1/2 d outside of the Kingdom), sent by International Money Order, payable to the Comptroller General, to the Patent Office, Sale Branch 2, Southampton Buildings London, W. C. 2. Danish, Kr 3 for each 5 pages sent to the Patentkommissionen Pøstboksvej 14, Copenhagen. Dutch, Fl 0 60, sent to Octrooiraad Oostdijklaan 2 The Hague, French, 5 franc plus postage, sent to l'Imprimerie Nationale, 27 rue de la Convention, Paris (15). German, R. M. 1, sent to the Reichspatentamt, Gitschiner Strasse 97-103, Berlin S. W. 61. Japanese, 5 sen plus postage sent to Hatsumei Kyokwai (Invention Association), 1, Yuraku Cho Ichome, Kojima-ku Tokyo. Norwegian, Kr 1 50 plus postage (usually Kr 0 30), sent to the Office de la Propriété Industrielle de Norvège, Middelthunegate 17, Oslo. Russian, 25 cents sent to Committee on Inventions of the U. S. S. R., Leningrad, 11, U. S. S. R., Swedish Kr 1 00 sent to Kungl. Patent- och Registreringsverket, Stockholm. Swiss, F. 1 20 per copy of 10 pages or less, F. 2 40, 11 to 20 pages, F. 3 60, 20-30 pages, etc., sent to Bureau fédéral de la propriété intellectuelle, Berne. Of Canadian patents manuscript copies only are obtainable. They are furnished by the Commissioner of Patents Ottawa. Estimates of cost can be obtained in advance. Belgian patents are not available in printed form.

AUTHOR INDEX

[An asterisk (*) indicates that the reference is to an abstract of a patent]

- Abeles I, 2759
Aberhalden F, 2693
2741, 2742 2763
*Abramson A, 2840
*Accumulatoren Fabrik A. G., 2648
*Aceta C m b H, 2852
*Acusa Rayon Corp, 2850
Adair G S, 2623
2747
Adams E Q, 2601
Adams R, 2711
Adams W W, 2852
Adkins H, 2685 2696
2688 2694 2713
2769
*Adler & Hentzen Maschinenfabrik, 2602
Adova A N, 2745
Alanaenko S S, 2872
Agati J A, 2871
*Akt Ges. Brown Boveri & Cie, 2604
2650 2681 2797
*Aktiebolaget B A Hjorth & Co, 2840
*Aktiebolaget Separator, 2806
*Aktiebolaget Separator Nobel, 2602
*Alberta W, 2679
Alewyn W F, 2673
*Alexander W, 2602
Ali M, 2791
Ali S B, 2808
Allard J, 2633
*Allardice T B, 2604
*Allégre J A, 2782
Allen H, 2639
Allen H S, 2643
*Allgemeine Elektrische Ges., 2604
2650 2651 2785
Allison B D, 2753
Allmand A J, 2616
Almelt C H, 2601
*Alsa Soc. smou, 2850
Alvarez E, 2804
Amfert P, 2740
*American Cyanamid Co, 2866
*American Smelting & Refining Co, 2785
Anderson A, 2780
Anderson B G, 2771
*Anderson E A, 2681
Anderson W, 2638
*Andrews L, 2605
*Asala American Corp of South Africa Ltd, 2677
*Asoda Rubber Co Ltd, 2676
*Apeldorn C S, 2792
*Appar F A, 2844
Arany S, 2793 2795
*Arias H, 2733
Aris K, 2630
Antake S, 2804
van Arkel C G, 2808
Arkibovich A. G., 2872
*Armbrust, C. W., 2604
*Armstrong, R., 2822
*Armstrong Fertilizer Works, 2802
Armstrong, E F, 2853
Arnot F L, 2538
*Arnouts, G., 2825
*Artex Chemische Fabrik G m b H., 2832
Asada R, 2867
Ash S H, 2833
*Ashenasy, P, 2604
*Asociación de productores de Yodo de Chile, 2649
Aston, J G, 2728
*Atlantic Coast Fisheries Co. The, 2781
Auer H, 2608
Aufhäuser N, 2852
*Auld S J M, 2844
Austin P R, 2688
*Autogen Gasacumulator A. G., 2837
v. Auwers K, 2710
Ayabe S, 2838
Aya S, 2805
Baas-Decking L. G. M, 2816
Bach, H., 2790
Bachmann, W. E., 2701
Bacher, H. J., 2691
Badanes, B. B., 2760
2784
Badger, W. L., 2785
Badstuber, W., 2721
*Baerwald F, 2649
*Bagby, R. B., 2781
Bagnoli E, 2794
Bailey, D. H., 2776
Bailey, K. C., 2632
Bain, D, 2665
Bairstow H, 2664
*Bakeite Corp, 2822, 2823, 2806 2867
Balabanov V, 2674
Balareff D, 2614
Balke M, 2620
Ball A M, 2852
*Ballmer, R. L., 2838
Baerje, D, 2639
Bangert, P, 2784
*Bangham P F, 2859
Banthal E J, 2747
Baranova R. L., 2632
Barash M, 2835
Barbato, L., 2718
Barber, C, 2799
Barbour O, 2768
*Bardgett W. E., 2650
*Barker, W. T., 2826
Barnes W H, 2639
Baron A., 2687

- *Barr, W., 2650
Barr, W. E., 2653
de Barros D., 2607
Bartels W., 2676
Barth, H., 2784
Bortholomew R. F., 2795
Bary, P., 2622
Bates, L. F., 2633
Boudisch O., 2657
*Bauer C., 2661
Bauer, K. H., 2665
Bauruschmidt A. J. Jr., 2640
Baum, K., 2633
*Baumgarten P., 2522
Bavink, B., 2635
Baxter, G. P., 2659
Bayerd, K. O., 2673
Bayles J. R., 2783
*Beardsley E. W., 2644
Beaver, J. J., 2631
Beccan, E., 2744
*Becher, W., 2520
Becker, E., 2676 2535
Becker, J., 2707
*Becker, J., 2640
*Beckers, A., 2639
Beckers, M., 2613
Beckers, J., 2610
Beche, C. C., 2601
*Beche, P., 2678
Beedham C. C., 2791
*Beermann, H., 2678
Beger, F., 2789
Behrens, B., 2610
*Beldmann A., 2603
Bell, P., 2726
Bell, R. W., 2617
Benedek, L., 2779, 2810
Benedict, P. O., 2762
*Bentley & Jackson, Ltd., 2651
*Berg, O., 2661
Berg, W., 2662
Bergami, G., 2766
von Bergen, W., 2654
Berger, F., 2608
Bergmann, E., 2611
Bergstrom, F. W., 2628
*Berio, D. W., 2638
Berlin, L. E., 2799
Bernadt, K., 2630
Bernoulli, E., 2613
Bernstein, A., 2736
Bersa, E., 2755
Bertram, G., 2764
Bertrano, G., 2756
*Bersebus' Metallhütte, G. m. b. H., 2678
Bessermans, A., 2763
Beutel, E., 2642
Bezzi, S., 2696
Bhate, S. R., 2780
Bhatnagar, S. S., 2605
Biancaleoni, G., 2744
Biedrzycki, B., 2624, 2627
Bigs R. P., 2719
*Billings, H. J., 2825
*Billwiler, J., 2849
Bilto, W., 2603
Binz, A., 2725
*Bino, A., 2814
*Birmingham Aluminium Casting (1903) Co., Ltd., 2679
Bissel, D. M., 2706
Burr, P. J., 2626
Bishop G. M., 2654
*Blampain, M., 2625
*Blank, A. D., 2784
Blank A. J., 2828 2629
Blau, M., 2637
*Blau Knox Co., 2628
Bleloch, W., 2640
Blennemann, E., 2667
Blom, A. V., 2615
Blom, J., 2603
Blörcher, H., 2783
*Blumenberg H., Jr., 2821
Blumberger J. S. P., 2653
Blumendel, H. B., 2657
*Blumenfeld, J., 2644
Blumrothel, B., 2632
*Blythe, G. E. K., 2604
*Blythen, S., 2766
*Boardman, C. C., 2659
Bobert, B., 2636
Boed, J., 2600
Bodendorf, K., 2695
Bodenstein, M., 2663
Bodnar, J., 2610
*Böhme, H. Th. Chem. Fabrik, 2624
*Böhmische Handels-
gesellschaft, 2633
*Bochlinger, C. F. &
Soehne G. m. b. H.,
2602, 2678, 2739
Böeken, J., 2606
Bogert, M. T., 2705
Bogschne, S. T., 2799
*Bohm, L. C. S. G.,
2680
Bokor, R., 2797
Boldagh, W. H., 2639
Bolton, R. P., 2751
*Bond, G. D., 2650
Bonuccelli, G., 2796
Boord, C. E., 2654
*Booth, V., 2637
*Booth, W., 2637
*Booth, W. J., 2637
Boothby, W. W., 2762
*Booy, J., 2833
Borchardt, W., 2765
Borchers, H., 2659
*Borggraeve, P., 2686
Borgström, L. H., 2668
Born, S., 2789
*Borsig, A. G. m. b. H., 2787, 2873
*Bossert, T. W., 2680
*Bosshard A., 2661
Bottoms, R. R., 2835
Boubet, C., 2619
*Boulade, A., 2645
*Boulogne, R., 2873
Bourro, M. C., 2747
*Boutaric, A., 2613, 2619
Bowman J. J., 2781
Bowen, I. S., 2641
Boy, R., 2691
*Boyer, H. L., 2645
*Brading D. A., 2739
*Bradley L., 2651
Bradley R. S., 2613
Bragagnolo G., 2784
Brähler P. S., 2644
*Braschk, K. W., 2604
v. Brand, T., 2759
Brandenburg, H. R., 2629
Brandt - Beaumont & (Mme.) 2764
*Bratting K., 2644
Braun H. J., 2616
Brauns R., 2614
Brückner, R., 2604
Bredon, H., 2609
*Bregat, J. H., 2785
*Breisig, A., 2637
*Breitloff - Versuchswerk, G. m. b. H., 2637
Brevort, M. J., 2601 2607
Breyer, 2672
*Brich & Spehr, 2782
*Briquetting & Carbons Syndicate Ltd., 2637
Brucce H. T., 2689
*British Cast Iron Research Association Ltd., 2679
*British Celanese, Ltd., 2650, 2659 2664, 2667
*British Glue & Chemicals, Ltd., 2675
*British Ropes, Ltd., 2677
*Bruchel, W., 2782
Brude, W. R., 2715
*Brodflex, C. M., 2650
*Bronsder, W. B., 2661
*Brown, J. I., 2650
Brooks, S. C., 2744
*Brown, A. B., 2643
Brown, J. B., 2776
Brown, L. A., 2771
Brown, F. E., 2799
Brown, R. C., 2625
Brown, R. E., 2687
*Brown, W. W., 2651
Brucer, E. M., 2842
Bruce, W. F., 2712
Brüder, P., 2605
van der Brugge, A., 2634
Brown, P., 2613
Brunel, A., 2755
*Bruno Stange, G. m. b. H., 2654
*Brutekus M., 2836, 2844
Bruss, H., 2634
*Bue H. R., 2733
Bucherer, H. T., 2629
Buck J. S., 2713
Burgin A., 2645
*Buffalo Electric Furnace Corp., 2785
*Buhlmann A., 2604
Buijck G., 2793
Bujarvics C., 2599
Bukoski H., 2659
Bull H. B., 2621
Bull L., 2622
Bullard E. C., 2636
Bumm F., 2764
*Burrak M., 2670
Burdakov B. A., 2871
Burger A. M., 2605
Burger H. C., 2607
*Burmeister H., 2625
Burr M. S., 2659
*Burrage A. C., Jr., 2676
Burrage L. J., 2616
*Burrows, H. O., 2679
*Buse, A. G., 2603
Buse S. A., 2700
Bussmeyer, H., 2631
Butcher C. H., 2616
Butler, B. S., 2664
Butler, J. A. V., 2625
Byers, B. M., 2770
Byers H. G., 2793
*Byers, J. A., 2679
Cabre, H., 2609
Cadwell F. H., 2634
Cady, W. H., 2654
Cahen, R., 2768
Caba, R. S., 2732
*Calrott, W. S., 2631
Caldwell, P. S., 2642
Calngart, G., 2622
Cameron, A. T., 2741
Campbell, A. N., 2608
Campbell T. P., 2646
Capel, W. H., 2609
Carlberg, H., 2670
*Carlton, R. A., 2650
*Caro, N., 2817
Carobbi, G., 2614
*Carpenter, I. C., 2844
Carpenter, T. M., 2732
Carr, F. W., 2646
Carnère, R., 2630
Carter, C. W., 2760
Cartwright, C. H., 2599
Casale, L., 2805
*Case, E. R., 2604
Cattrian, E., 2662
Caughley, F. G., 2874
*Cave Brown Cave, T. R., 2846
*Cellulose Corp., 2627
*Central - Europäische Schwimm - Aufbereitungs - A. G., 2677
*Centrifugal Castings, Ltd., 2679
Chakravarty, K., 2835

- Chalkley, H. W., 2770
 Chamot, E. M., 2635
 Chance Bros. & Co., Ltd., 2827
 Chapin, R. M., 2657
 Chapman, A. W., 2790
 Chapman, J., 2701
 Charles, 2666
 Charles A. P., 2741
 *Charles, E., 2738 2769
 Charrin V., 2670
 *Châtillon (Società anonima italiana per la seta artificiale), 2849
 Cheshbrough E. W., 2663
 *Chem. Fab. von Heyden, A.-G., 2734
 *Chem. Fab. vorm. Sandoz 2662
 *Chemieprodukte, G. m. b. H., 2631
 *Chemieverfahren, G. m. b. H., 2318
 *Chemische Fabrik von Heyden, A.-G., 2525
 *Chemische Fabriken Karl Albert, G. m. b. H., 2866
 *Chemische Fabrik Ludwig Meyer, 2804
 *Chemische Fabrik Schonenwerd H. Erlinger, A.-G., 2865
 *Chemisches Laboratorium für Toxikologie und Tox. Industrie-Zeitung H. Seyer & E. Cramer G. m. b. H., 2786
 Chertov, S. I., 2630
 Chichibabin, A. E. 2725, 2727, 2729
 Chick, H., 2750 2760
 Christensen, E., 2830
 Christiansen, W. G. 2813
 Christman G. 2762
 Chrzanoski A. 2801
 Churchill J. B. 2784
 Chwala A. 2600
 Cim V. 2796
 van Cittert P. H. 2607
 Claassen H. 2872
 Clapp W. B. 2751
 *Clark P. L. 2649
 Claussen 2796
 *Clavel R. 2861
 Clegg E. L. G. 2665
 Clementi A., 2744
 Coblyn J. H. 2635
 Coffman V. 2743
 Cohen R. 2751
 Cohen, E. 2613
 Coburn-de Meester W. A. T. 2613
 Cole, H. I. 2600
 *Cole W. H. 2691
 Colin Russ A. 2873
 *Collet, H., 2603
 Collins, L. H., Jr., 2750
 Collins W. D., 2659
 *Colloid - Chemische Forschungen, A.-G. 2419
 Colobran V., 2772
 *Colomb H. 2650
 Colombari, L., 2659
 *Colville D. & Sons, Ltd., 2650
 *Commercial Solvents Corp., 2735 2806
 *Compteur des textiles artificiels (Soc. anon.) 2650
 *Compagnia Metallurgica de Marzotto 2465
 *Compagnie continentale pour la fabrication des compresseurs et autres appareils 2437
 *Compagnie des mines de Vicoigne Nucua et Drocont 2639
 *Compagnie française d'accumulateurs électriques 2644
 *Compagnie française de l'Inde et de l'Alger 2824
 *Compagnie française pour l'exploitation des procédés Thomson Houston, 2680 2787, 2866
 *Compagnie générale d'électricité 2423
 *Compagnie Lorraine de charbon pour l'électricité 2644 2865
 *Compagnie nationale de matières colorantes et manufactures de produits chimiques du nord réunis établisments Kuhlmann 2822 2837 2850
 Conant J. B. 2733
 *Condrup C. O. 2839
 *Conlan E. J. 2827
 Connor R. 2694
 *Consortium für elektrotechnische Industrie, G. m. b. H. 2738
 Cook J. W. 2716
 Coolhaas C. 2411
 Cooper H. M. 2833
 Cooper W. L. 2719
 *Cooper A. C. 2709
 Copeman P. R. & Co. 2779
 Copisarow M. 2822
 Copley M. J. 2600
 Coppens A. M. 2759
 Copson R. L. 2601
 Corbellini A. 2716
 Cora, C. B. 2660
 *Cornic V. 2443
 *Cornillie A. L. P., 2406
 Cotton P. H., 2476
 de Coulson, A., 2761
 *Coupau M. G. 2406
 *Courtaulds Ltd., 2418, 2450
 Courtney H. 2863
 Cowell, G. R., 2761, 2762
 Cox R. F. B. 2646
 Cramer H. 2713
 Cramer H. L., 2694
 Crandall F. K., 2690
 Crane, E. V., 2674
 Craven F. C., 2441
 *Crawford R. A., 2570
 *Crispien A. R., 2438
 *Cress C. P., 2786
 Cross W. R., 2870
 Crow, A. D., 2612
 Cucky, J., 2791
 Cuyler Z., 2407
 Caillard J., 2753 2775 2776
 Caenel F., 2662
 Cunningham G. E. 2614
 Cunningham J. 2432
 Cunniff L., 2749
 Curli A. L., 2699
 Czab E. 2606
 Daddow W. T. 2701
 *Dagand J. M., 2782
 Dakers J., 2610
 v. Dalwitz Wagner R., 2634
 Dalvi P. H. 2874
 Damerell V. R., 2644
 Damsch J. 2833
 Dannenberg H. 2405
 Dammeyer F. 2759
 *Darling F. R. 2832
 *Darrah W. A. 2746
 Das R. 2857
 David L. 2809
 Davidow J. 2467
 Davidson A. W. 2624
 Davies W. 2702
 Davis A. C. 2424 2829
 *Davis C. A. 2679
 *Davis F. 2860
 *Davy W. J. 2603
 Dawson, T. R. 2875
 Dean A. S. 2791
 Dean R. S. 2672
 *Deberth H. 2650
 *Derlone G. L. 2861
 Dedinsky G., 2775
 *Detekamp C., 2792
 *Detp L. 2602
 Del Trodon A. 2797
 *Deon Delwig C. 2650
 Demay A. 2668
 Demfay L. 2703 2704
 Denung H. G. 2606
 Denung W. R. 2612
 Demolen A. 2794
 Demus E. 2608
 Denny, P. E., 2756
 DePottier, P., 2764
 Derge, G. J., 2689
 Derksen J. C., 2623, 2846
 Derry, D. R., 2669
 Desai R. V., 2747
 *Desmets, P., 2814
 *Desmoulin, H. A. 2792
 Desmoulin P., 2747
 Desvergues L., 2699
 Deterre J. 2400
 Deulofen V., 2853
 *Deutsche Abwasser-Reinigungs-Ges. m. b. H. 2792
 *Deutsche Babcock & Wilcox, Dampfkessel - Werke A.-G., 2602
 *Deutsche Gold- und Silber-Schneidmanufaktur vorm. Roessler 2782 2403 2418, 2417 2828
 *Deutsche Hydrierwerke A. G., 2405 2478
 *Deutsche Schmelz- und Raffineriewerke A. G., 2420
 *Devant, M., 2755
 Desautel S. 2766
 Dhar N. R., 2843
 *Diamco A.-G., 2651
 *Diamco A. G. 101
 Glühlicht 2648
 Dickens P. 2600
 Dickinson H. C., 2841
 Dieckner K., 2637
 Dieckmann C., 2647
 Dieckmann W. J. 2731
 Diemke J. W. 2699
 *Diethelm H. 2839
 Dietrich L. R. 2832
 Dietrich W. 2800
 Di Giéna, J., 2791 2795
 van Dyk J. A. 2848
 Dingemans E. 2731
 *Dinley C. F., 2824
 Ditmar R. 2785
 Dittmer M. 2469
 Dodd R. 2776
 Döhrmann V., 2454
 Dörner, L. 2780
 Doerner H. O., 2674
 *Doherty, H. L., 2837
 *Doherty Research Co., 2619 2631
 Dole M., 2627
 Dombitsky C., 2870
 Donath W. F., 2779
 Doniger M., 2463
 Dorsch, K. E., 2825
 Doughty, H. W., 2689
 *Dovel J. P. 2603
 *Dow Chemical Co. 2450
 Downs H. C. 2670
 Downes J. R. 2790
 Downs W. G., 2764

- Dos, A. W. 2697
 *Drew, R. B. 2575
 Deems 2505
 Drewyer, G. P. 2749
 Dreyfus C. 2560
 2862
 *Dreyfus H., 2735
 2849, 2850 2542
 Druce J. G. P. 2608
 Dubinin M. M. 2616
 Dubois L. 2512
 Dubsky, J. V. 2653
 2660 2708
 *Ducamp A. J. 2545
 2873
 Ducceschi, V. 2619
 Duckham, A., 2511
 *Duckham A. M.,
 2810
 *Duckham A. McD
 2603 2539
 Ducas E. H. 2664
 Dulrasse C. 2531
 Dumanski, A. V. 2570
 *Dumas M. G. 2679
 Dunsar, R. E. 2606
 Dunlavy, H. 2793
 *Dunlop Rubber Co
 Ltd., The 2531,
 2576, 2577 2878
 Dusa, J. A., 2644
 Dunnewild T. J. 2793
 Dunschiff, H. B. 2656
 Dunsheath P., 2782
 Dunstan, P. T., 2673
 Dupont, A., 2609
 *Dupre, A. F. H.,
 2840
 *Duplan, F. P. F. S.,
 2544
 Dupont, G., 2633
 *Du Pont de Nemours,
 E. I., & Co., 2875
 *Durand & Huguenin,
 S. A., 2859
 *Durand, J. B., 2679
 Dulce, W. G., 2553,
 2874
 Durrans, T. H., 2543
 Dworak, L., 2795,
 2797, 2799
 D'yachkovskii, S. I.,
 2619
 *Dynamit-Actien-Ges.
 vorm Alfred Nobel
 & Co., 2739, 2740,
 2853
 Eagleson, W. V., 2815
 Eaton, F. M., 2794
 Eberhard H. M., 2785
 Ebert, P., 2633
 Ebert, L., 2696
 Escha, A., 2833
 Eckstein, H., 2663
 Eddy, C. W., 2599
 Eddy, W. H., 2760
 Edén, B., 2640
 *Edmonds, W. J.,
 2739
 Effront, J., 2697
 Efimov, M. G., 2870
 Ege, R., 2746
 Egeberg, B., 2646
 Eggert, J., 2643
 Pgorov I. I., 2570
 Pgorov O. I., 2708
 *Eggenh. landlämpa
 & Villamowdg. R. &
 avfnyttarsk. 2603
 2649 2651
 Ehrenhaft F. 2643
 von Ehrenhaft F.
 2741 2742
 *Eichenberg G., 2670
 *Eichenberger F.
 2735
 Eide A. P., 2698
 Eilhorn G. E., 2573
 Eilstein O., 2659
 Eisenbrand J., 2609
 *Eisengewerl. Mas.
 schinen und Pappm.
 fabriek F. A. Möns
 ner G. m. b. H. 2551
 Ekst W. 2672
 Ekholm I., 2544
 Eklin P., 2656
 Eklin W. 2754
 *Elkington V., 2506
 *Elmer W. S., 2792
 Ellis G. D. 2637
 Ellis J. W. 2641
 *Ellis P. H., 2544
 Elmquist R., 2654
 Emle H., 2606
 Emmermann G., 2652
 Enders I., 2631
 Endres G., 2750 2765
 Engel L., 2611
 *Engelstad A., 2786
 *Engelstam, G., 2794
 *Engelstam, S., 2533 2842
 *Engelstam, T., 2641
 *Engelstam, J., 2781, 2784
 *Engelstam, L. J., 2644
 Ernst, R. C., 2843
 van Erp, H., 2706
 *Ernst, H., 2843
 *Eschbach, W., 2853
 Eschmann, W., 2874
 Escudero, A., 2749
 *Essling, F., 2845
 Ess, T. J., 2675 2534
 *Establissemens F.
 Orange et Cie., 2877
 von Euler, H., 2744,
 2764
 *Eutis, P. A., 2781
 Evans, E. J., 2649
 Evans, F. R., 2775
 Evans, G. L., 2833
 Evans, W. V., 2624
 Evenden, P. F., 2647
 Pfing, W. W., 2616
 *Fabriques de produits
 de chaux organique
 de Laire, 2822
 Fachtel, G., 2786
 *Faerber, F., 2806
 Faerman, S. H., 2714
 Faerman, G. P., 2858
 Fairbourn, A., 2692
 Fairbrother, P., 2620
 Fairbrother, R. W.,
 2767
 Fairchild, J. G., 2662
 Fairhall L. T., 2661
 Faltin, A., 2775
 *Fancy, J. G., 2524
 Fanshawe J. R., 2536
 Farhas, H., 2529
 Faeham P. C., 2677
 Farr H. V., 2659
 Faust O., 2547
 *Fauch P. L., 2670
 Fayed A. R., 2774
 *Fayer C., 2603
 *Fedeler J. H., 2792
 Fehér D., 2793 2797
 Fehrbach K., 2764
 Fehrmann R., 2535
 Feit, W., 2604
 Feigenthaler W.
 2600
 Felchani A., 2796
 *Fellner & Zeigler
 A. G., 2603
 *Feltin & Guillaume
 A. G., 2605
 Fernelius W. C., 2629
 2629
 Ferner G. W., 2660
 Ferschl C., 2794
 de Fertilis J. Y.
 2796
 French V. H., 2632
 Fichter, P., 2615
 Fidge F. L., 2606
 Fiedler A. G., 2533
 Fildorfer, M. S., 2873
 Finkey, J., 2533 2534
 Finnerty A. P., 2752
 Fissi C., 2711
 *Firman, W. G., 2605
 Fischbeck, H. J., 2644
 Fischbeck, K., 2656
 Fischer, 2808
 Fischer, P. P., 2766
 *Fischer, R., 2654
 Fischer, V., 2635
 Fischer, W., 2733
 Fischler, F., 2772
 Fish, P., 2771
 Fisher, F. D., 2606
 Fitch, J. B., 2781
 Fischen, M. A. V., 2760
 Flannery, M., 2663
 Fleischner, I., 2772
 *Flemming, P., 2602
 Fletcher, G. L., 2635
 Fletcher, J. F., 2674
 *Fletcher J. P., 2879
 Fleury, P., 2749
 *Fledner, A., 2601
 Fledner, L. J., 2606
 Fodman, E. D., 2620
 Fodor, K., 2780
 Folkers, K., 2686,
 2694
 *Follet, A., 2678
 Fontaine, 2790
 Foobey, W. L., 2683,
 2692
 Forbes, J. C., 2751
 Forrest, S. D., 2665
 Forsén I., 2829
 *Forster, A. L., 2827
 Fortson, N. S.,
 2644
 Foshler, R. J., 2747,
 *Fouché M. F., 2527
 Fount C. G., 2656
 Fowweather F. S.,
 2766
 Fox C. S., 2515
 *Fox O., 2603
 *Frankl S., 2515
 *Frankl M., 2539
 Frahm P. D. G., 2799
 *Franko W., 2540
 2542
 *Frank A. R., 2517
 Fraps G. S., 2762
 Frarlini B., 2766
 *Friedenhagen K.,
 2740
 Friedborn S. B., 2775
 *Freedman P., 2603
 Freeman, J. V., 2659
 Freiburg A., 2549
 *Freilichen A., 2502
 *Freinkel E., 2544
 Freud J., 2751
 Fries J. H., 2612
 *Friedrich W., 2531
 Friend J. N., 2658
 Friese P. W., 2644
 *Frigliani Corp.
 2787
 Froehner M. Jr.
 2791
 *Frost D. P., 2577
 Forth R., 2611
 Fuller J. P., 2794
 Fuller R. F., 2671
 Fulton C. G., 2645
 Fulton H. R., 2781
 Furman N. H., 2659
 Furnas, C. G., 2516
 *Gade M., 2506
 Gahrner, K., 2834
 Galea, V., 2730
 Galhber, B. W., 2664
 Galhbar, D., 2761
 Galhbar, Z., 2834
 Ganguli R., 2634
 Gannon, E. N., 2629
 *Gardel, G. L., 2786
 Gargill, S. L., 2751
 Garner, W., 2790
 Garrick, F. J., 2604
 Garraw, F. C., 2604
 Garstang W. L., 2631
 *Gaudet, J. V., 2845
 Gaxley W. C., 2774
 Gebrüder Klotz 2873
 *Gebr. Wansleben,
 2860
 Geck, W. H., 2833
 Geering M. C., 2721
 Gebie, H., 2600
 Geier, H., 2670
 Geiger, A., 2760, 2763
 Galhorn, B., 2766
 *General Technical
 Co., Ltd., 2845
 Genin, G., 2817
 *Gentil, P., 2826
 *Gerber, N. Cu m.
 b. H., 2782
 *Gerber P. T., 2860
 Gerecs, A., 2732
 Germuth, P. C., 2661,

- Gerry L. S. 2452
 Gerstenberger D. P. 2842
 Gerstner F. 2662
 *Ges. für Kohlentech. mit u. H. 2534
 *Ges. für Landw. Maschinen A-G 2435
 *Gewerkschaft Victor Stickstoffwerke 2420
 *Gewerkschaft Wallram 2679
 Gez. M. 2604
 *Gfeller H. 2512
 Ghini E. 221
 Ghuron D. 2645
 Ghosh J. C. 2533
 *Gibbons W. A. 257
 Gibbs R. S. 2661
 Gibbs R. S. 2661
 Gibson G. P. 2692
 Gibson R. E. 2626
 Gickhorn J. 2750
 Gidró J. 2793
 *Giebmann A. 2544
 Giese A. C. 2744
 Giese R. J. 2711
 *Giehrst, R. W. 2649
 Gill R. W. R. 2536
 *Gillespie H. R. 2788
 Gilman, H. 2657
 Ginsburg R. B. 2631
 Gläthe L. 2800
 *Gläthe J. 2831
 Gläthe, E. S. 2932
 Glocker, R. 2639
 *Gloess, P. 2557
 *Glover, W. H. 2850
 *Glockmann E. 2811
 Glud W. 2647
 Gnadinger, C. N. 2401
 Gnessin, Y. D. 2811
 *Godel, A. 2869
 Gombey, S. 2772
 Goettlich M. 2761
 Goets, C. 2672
 Gohring C. F. 2869
 Goldfarb, Y. L. 2719
 *Goldmann H. 2649
 Goldstein L. 2608
 Goldsworthy L. J. 2697
 Coll, G. 2628
 Gonzalez, V. F. 2675
 González de la Vega M. 2462
 *Goodall C. 2532
 Goodman J. B. 2623
 *Goodwin R. K. 2828
 Goranson R. W. 2612
 *Gorham W. G. 2577
 Gorter C. J. 2620
 Gortner R. A. 2621
 *Goskar T. A. 2637
 Goudry R. F. 2780
 Gräven H. 2661
 Graf S. H. 2743
 Crafe F. 2764
 Grafe V. 2785
 *Granges L. 2406
 *de Granville, J. 2624
 Grasser, G. 2573
 Gratz, O. 2776 2777
 *Gravell J. H. 2641
 Gray P. W. 2610
 Gray, I. 2747
 *Gray T. H. 2423
 *Credit, P. 267
 Green J. 2773
 *Greenwalt, J. P. 2531
 Greenfield H. 2717
 *Greenhill M. 242
 Greenstein J. P. 2627
 Greer, L. 2669
 Grewe E. 2771
 Griffiths, J. G. A. 2641
 Griffiths, W. J. 2767
 Grignard V. 2762
 Grunne, C. 2409
 Grimshaw, W. R. 2612
 Grinberg A. A. 2654
 Grindrod G. 2761
 Grellman A. 2750
 *Gross, J. E. 2603
 Grove & Palmer, F. 2451
 Gröndler, W. 2667
 *Gruenwald T. 2673
 *Grühl, P. 2439
 *Grunert & Guarnetti 2450
 Cryder, D. O. 2634
 Gurner F. 2405
 Guéhen C. 2616
 *Guénot, C. P. A. 2675
 Guerrant, N. B. 2763
 Guertler, W. 2632
 Guba, S. K. 2715
 Guilbert G. 2648
 Guillemin C. O. 2753
 Guillet L. 2676
 *Guise C. M. M. 2570
 *Gull A. E. 2450
 Gulyás A. 2401
 Gump W. 2730
 Gums W. 2536
 Gupta J. D. 2638
 Gustavson K. H. 2571
 *Gutehoffnungsbütte Oberhausen A. G. 2766 2638
 Guxson G. 2663
 Gyzander A. K. 2654
 *Haas F. 2432
 *Haas F. J. 2604
 de Haas W. J. 2609 2610
 Haase E. 2698
 Hackh I. W. D. 2733
 Hadert H. 2464
 Haebler H. 2634
 Hägg G. 2615
 *Hagerl H. 2602
 *Haglund T. R. 2649 2619
 Haala F. 2667
 Hallday R. G. 2774
 Hamad M. A. 265
 Hamilton W. F. 2740
 Hammerich T. 2710
 *Hammer, R. P. F. A. 2437
 Hanson, J. 2671
 *Hamp J. 2755
 Hanamura, S. 2404
 *Händler, R. 2614
 Handozky H. 2769
 Haskin, S. 2431
 *Hannova Chemical and Manufacturing Co. 2649
 Hansen A. A. 2402
 Hansson T. 2643
 Hapgood F. C. 270
 Hara, R. 2606
 Harada, T. 2660
 Harding, E. R. 2776
 *Hardy, H. 2439
 Hardy, W. 2618
 Harig G. 2632
 Harluff, J. H. 2612
 Harkins, W. D. 2617
 Harland J. S. 2665
 Harris R. H. 2774
 *Harrison, W. 2448
 Harrow R. 2635
 Hart, M. C. 2733
 Hart, R. 2468
 Hartley H. 2625 2626
 Hartmann, B. 2604
 *Hartmann, M. 2604
 *Hartmann, R. A. (Pirma) 2403
 Harvey-Gibson R. J. 2635
 *Harzer Maschinenbau, G. m. b. H. 2440
 Haas K. H. 2780
 *Hawthorn A. 2406
 Hawthorn A. 2754
 Haxton H. D. 2751
 Ham H. B. 2649
 Hlawek O. 2698
 *Hatfield H. S. 2603
 Hates, G. 2627 2628 2794
 Hatzfeld, K. 2633
 *Hauhold C. G. A. G. 2431
 *Haug J. S. 2535 2639
 Hauss F. 2754
 Hawley F. G. 2613
 Haynes, L. 2755
 Haskins E. O. 2719
 Haxby V. 2760
 Hedberg H. D. 2440
 *Heerdt - Langer G. m. b. H. 2517 2825
 Hegel, K. T. 2635
 Heilmann L. V. 2771
 Heisinger R. 2669
 Heilmeyer L. 2745 2782
 Heintz E. 2646
 Heiser A. 2874 2875
 Heitz W. 2722
 Held N. A. 261
 Held S. 2606
 Heifenstein, A. 2733
 *Heifrich B. 2740
 Heller G. 2709 2711
 *Hene, R. 2614
 *Henkel & Co. G. m. b. H. 2624 2670
 *Henkel & Knaus Chem. Fab. G. m. b. H. 2675
 Henry, T. A. 2730
 Hentze G. 2780
 Hepburn, J. S. 2765
 Herporth T. C. 2444
 *Heraeus W. C. G. m. b. H. 2630
 Herbert F. K. 2747
 Heed C. W. 2773
 Heitler, J. 2661
 Herman F. H. 2400
 Herndon T. C. 2601
 *Herring P. H. 2844
 Herman R. 2796
 Hermann, R. 2600
 Hertel P. 2430
 *Herbst, E. C. 2443
 Herring L. 2404
 Hervey W. 2665 2410
 Hira, K. 2623
 Hira, W. C. 2710
 Hesse, K. 2765
 Hesse P. 2411
 v. Hesser, P. V. 2436
 Hesser P. 2433
 *Hesse C. 242
 Hey, W. 2636
 Heyes T. F. 2433
 Heyl P. W. 2733
 *Heyl G. E. 2427
 Heyne G. 2661
 *Heynemann, H. 2625
 Heyrovsky J. 2679
 Hiedemann F. 2600
 Higashi K. 2606
 *Hilger G. 2403 2639
 Hill H. G. 2577
 Hill, M. 2765
 *Hiller H. 2619
 Hinselwood C. N. 2531
 *Hirsch Kupfer- und Messingwerke, A. G. 2679
 Hirsch W. C. 2448
 *Hirschler R. 2747
 Hirschfelder P. 2672
 Hirschberg R. 2609
 Hochrein M. 2765
 Hodges F. 2674
 *Höfken M. 2449
 *Hönnicke G. 2654
 *Hoesch & Co. 2448
 Hoffmann J. 2634
 *Hoffmann La Roche F. & Co. A-G 2414 2625
 Hofman I. L. 2799
 Hoffmeyer R. 2750
 *Holmes L. C. 2465 2736
 *Holschneider F. 2759
 Holts F. 2759

- *Holtz, J. C., 2535
 *Holtzappel, A. C., 2544
 *Holzverkohlungs Industrie, A. G., 2735
 2506
 Hong, P., 2573
 *Honnay, C., 2537
 Hoogeveen, A. P. J., 2715
 *Hookey, L. J., 2539
 Hoover, C. F., 2789
 *Hopkinson, E., 2577
 *Hopkinson, H., 2562
 *Hoppe, O., 2604
 Hoppmann, 2809
 Hornemann, T., 2606
 Horner, J. G., 2676
 Horner, W. W., 2790
 Horovic, A., 2672
 Horst, C., 2516
 Horváth, I., 2779
 2780, 2810
 Houben, J., 2733
 Hough, C. T., 2748
 *How, H. W., 2602
 Howard, R. S., 2606
 Howland, R. B., 2746
 Hsueh, C. Y., 2650
 Hubbard, R. S., 2751
 Hubba, C. L., 2770
 Hudson, R. F., 2662
 *Hulschmidt, C., 2604
 Hughes, O. L., 2626
 Huma, Rotbery, W., 2639
 Hummer, O., 2504
 *Humphreys & Glasgow, Ltd., 2535
 Hunter, C., 2747
 Hunter, W. H., 2703
 Huppmann, 2808
 Hurd, C. D., 2683
 Hutchison, W. K., 2629
 Ibañez, R. F., 2745
 Ide, W. S., 2713
 Ida, S., 2638, 2639
 *I. G. Farbenindustrie, A. G., 2602, 2605, 2648, 2650, 2651, 2653, 2654, 2677, 2678, 2680, 2681, 2733, 2734, 2735, 2736, 2737, 2738, 2739, 2740, 2786, 2787, 2803, 2804, 2813, 2814, 2815, 2818, 2819, 2820, 2821, 2822, 2824, 2825, 2827, 2828, 2831, 2837, 2838, 2839, 2843, 2846, 2847, 2848, 2849, 2850, 2853, 2855, 2856, 2857, 2858, 2859, 2860, 2861, 2864, 2865, 2866, 2867, 2869, 2876, 2877, 2878.
 *Iltener, Glas Instrumenten - Fabrik Albert Zuckschwerdt, 2601
 *Imhoff, M., 2561
 *Imperial Chemical Industries, Ltd., 2605, 2610, 2736, 2737, 2739, 2740, 2821, 2847, 2857, 2858.
 *Impeya Holzimpfungs- und Holzverwertung, A. G., 2532
 Imre, L., 2616, 2618
 *Industriemiska Aktiefabriken, 2517
 *Ingersoll Rand Co., 2605
 International Gradina, A. G., 2781
 *International Industrial & Chemical Co. Ltd., 2520
 *International Metall, A. B., 2605
 *International Nickel Co., Inc., The, 2618
 *Internationale Gase, H. Hoff & Co., 2823
 Intonti, R., 2645
 Ionas, A. A., 2799
 *Irwin, H. G., 2603
 Istell, H. G., 2576
 Ishibashi, M., 2663
 Ishii, S., 2620
 Ishikawa, H., 2540
 Iwasaki, M., 2643
 Ivány, I., 2779, 2804, 2805
 Ivy, A. C., 2749
 *Jaacks-Ballester, F., 2570
 Jablonski, K., 2621
 Jackson, I. K., 2702
 Jackson, S., 2790
 Jacobs, F., 2875
 Jacobson, K., 2657
 *Jacoby, M., 2804
 Jacques, A. G., 2756
 Jaeger, W., 2670
 Jänecke, R., 2830
 *Jahn, A. R., 2602
 *Jakob, G., 2806
 Jakóby, J., 2835
 *Jander, G., 2819
 Janček, M., 2610
 Jankelson, I. R., 2751
 Jantson, H., 2798
 *Jaubert, G. F., 2651
 Jeffrey, F. H., 2632
 Jenkins, I., 2640
 Jenkins, T. A., 2761
 Jenka, H. N., 2767
 Jensen, E. C., 2652
 Jessop, G., 2615
 *Jessup, A. C., 2649
 Jeulroy, A., 2633
 Jilison, W. R., 2670
 *Jochem, O., 2667
 Joffe, A., 2608
 Joffe, I. S., 2707
 Joffe, V., 2609
 Johnson, A. H., 2773
 *Johnson, F., 2802
 2819
 *Johnson, F. A., 2607
 Johnson, G., 2658
 Johnson, K. A., 2834
 Johnson, R. H., 2783
 Johnson, T. B., 2708
 2722
 Johnson, W. T., 2775
 Johnston, W. S., 2782, 2834
 Johnston - Taylor, 2607
 Jols, H. S., 2868
 Jomay, W. F., 2629
 Jones, C. L., 2783
 Jones, D. B., 2763
 *Jones, F. A., 2878
 Jones, F. L., 2536, 2639
 Jones, G. W., 2612
 Jones, M. P., 2601
 Jones, P. C., 2878
 Jones, W. A., 2671
 Jones, W. J., 2702
 Jons, G., 2743
 *Jordan, H., 2740
 Jorri, 2706
 *Jost, W. F., 2833
 Jost, W., 2630
 Joubert, C. J., 2601
 Juel, C., 2770
 Judkins, H. F., 2776
 Jurist, A. E., 2813
 *K. D. P. Ltd., 2876
 Kadonaki, H., 2688
 *Kaiser, L., 2818
 Kalashnikov, R. Y., 2873
 *Kali Chemie A. G., 2602, 2640, 2670
 *Kali - Forschungsanstalt G. m. b. H., 2664
 *Kalle & Co., A. G., 2654
 *Kallung, B. M. S., 2650
 Kamada, T., 2626, 2628
 Kammerling, S. E., 2733
 *Kamps, H., 2630
 *Kamps, O., 2679, 2680
 *Kappeler, H., 2822
 Kappelmeier, C. P. A., 2729
 Karácsonyi, L., 2774
 Kargin, V. A., 2679
 Kerk, D., 2645
 Kernhan, M., 2760, 2761
 *Karpén, S. & Bros., 2740
 *Karrénberg, G., 2604
 Karter, F., 2733
 Karahan, M., 2761
 Karsten, P., 2869
 Kasural, T., 2743
 Katz, J. R., 2623, 2846
 Kaufmann, H., 2631, 2865
 Kaufmann, H. P., 2865
 *Kaufmann, L., 2737
 Kawamura, K., 2794
 Kaye, G., 2767
 Keane, J. C., 2871
 *Keasley, W. K., 2603
 Keeler, C. R., 2791
 *Keel, C., 2603
 Kehr, 2791
 Kehr, R. W., 2599
 Keil, W., 2691
 Keller, K., 2607
 Kelly, T. L., 2662
 Kennet, J., 2604
 Kenyon, G. H., 2833
 Kessler, N., 2753
 Kernot, J. C., 2874
 Kerley, Z. I., 2767
 Keyes, P. G., 2612
 *Klepe, H. R., 2650
 Kiewitbach, T. A., 2780
 Kiesel, O. R., 2833
 de Kiewit, T., 2710
 Kik, M. C., 2738
 Kilp, W., 2805
 Kingsley, J. F., 2788
 Kinnerley, H. W., 2760
 Kino, K., 2693
 Kinney, E. L., 2641
 Kirsh, W., 2798
 Kiskadden, F. L., 2899, 2903
 Kiss, A., 2631
 Kistler, A. S., 2620
 *Klausmann, H., 2640
 Klausvik, O., 2630
 Kleber, B., 2876
 Kleiderer, B. C., 2711
 Klein, A. J., 2747
 *Klein, F. G. C., 2818
 Klein, H., 2765, 2834
 Klein, W. I., 2789
 *Kleinberger & Co., A. G., 2831
 Klement, A., 2696
 *Klempsch, W., 2679
 Kling, D. H., 2747
 *Klopsteg, P. R., 2787
 Kloster, G., 2740
 *Knapp, A., 2677
 Knapp, O., 2626
 Knapp, W., 2676
 Kneuer, H. O., 2612
 *Knoll, A. G., Chem. Fab., 2815
 Knoops, 2644
 Kobayashi, K., 2540
 Kobeko, P. P., 2611
 Kober, S., 2599, 2751
 Kochnar, K., 2755
 Koch, F. K. V., 2616, 2627
 Kocheshev, K. A., 2702
 Kögel, G., 2601
 *König, Friedrich August - Hütte A. G., 2870
 Körber, F., 2643
 Koller, L., 2817
 Kohler, E. P., 2712
 Kohlshütter, H. W., 2635
 *Kohman, H. A., 2788

- *Kobara O & Co., 2550
 *Kobut S 2531
 Kok J A F 2552
 Kolkwa R 2589
 von Kolmitz, H. 2664
 Ko-pakow V I 2571
 Kothof I M 2625
 2625 2655
 Kon G A R 2695
 Konora, C 2772
 Kondo, H 2731
 Kondo M 2775
 Konek, P 2510
 Koneberg T 2747
 Koothas D R, 2511
 Kopp H 2635
 *Koppers, H 2540
 *Koppers, H A-G., 2651
 Koppová, A. 2794
 Kormann I M., 2783
 Kornfeld C 2641
 Korpácsy I., 2777
 Koryzev A. B (MfU.) 2632.
 Kos, J 2606
 Kosaka H, 2753
 Kottmann, L. 2600
 Kovats L. de T., 2797, 2900
 Kuyanagi K., 2529
 Kresler L 2673
 Kresmer A. J., 2541
 *Kramer C., 2603.
 Kraus N W., 2633.
 Krausku K. A., 2590
 Krata H, Jr., 2791
 Kratz E 2619
 Kramers, C. A., 2623
 Krampe, H. J., 2601
 Kraus, C. A., 2624
 *Kraus, C. S., 2521
 Kraus, W., 2576
 Kraus H., 2605
 Krebs W., 2752
 *Krebs & Co., G m b H, 2786 2821
 Krieger H 2783
 *Kristenmon A. 2520
 Krivonos F F 2690
 Krid A., 2605 2675
 Kröger C., 2634
 Krocak A 2780
 Kropp B 2754
 *Kroger, I A S., 2792
 *Krupp F A-G 2690 2651
 *Krupp F Grusonwerk A G 2677
 Kryl P 2572 2873
 Kubelka V 2843
 Kühles W 2792
 Kuhlmann G 2809
 Kuhn A 2509
 *Kuhn R 2735
 Kukuk, F 2670
 *Kunstharzfabrik F Pollak G m. b. H. 2822
 Kunze, P., 2640
 Kuputa, J 2776
 Kural, M., 2708
 Kurchak I. V., 2511
 Kurelec, V., 2751.
 Kurbanov D., 2698.
 Kurbanova, A. I 2727
 Kasharyev M A 2753
 Kasumina, N 2768.
 *Kattier H. L. 2531
 Kutzelnigg A., 2642
 Kaula O 2561.
 Kuwata, T., 2659
 Laflitte, P., 2532.
 Lagemann, A., 2621
 *Lager E 2602
 *Lahousseau, R., 2602.
 *Larson, A 2520
 Laird, H C 2689
 Laker, A. 2669
 Lambert 2795.
 *Lambert B., 2517
 Lambourn, H., 2659
 La Mer, V. K., 2626
 Lampe, B 2573
 Landers G., 2656
 Landova, M., 2542
 Lane, E. C., 2541
 *Lang, A., 2654
 *Lang C., 2654.
 Lange, H., 2609
 Langst, O., 2755.
 Langfeldt E., 2781
 Langmuir L., 2536
 Langton, H M., 2569
 *Langwell, H 2757
 Lank, J., 2794.
 *Lantz, R., 2653
 Lány, B., 2643
 La Porta, A., 2443
 Laquer, E., 2751.
 Laronov, W T., 2766
 Larmer R. K., 2774
 Larson R. L., 2507
 *Larson, A. T., 2821
 *Larson M., 2517
 *Lasky M. 2827
 *La Soudiere electrique autogene (Soc. anon.) 2691
 Lategas F N 2833.
 Lauer W M. 2696
 *Lauter, W M., 2815
 *Laws J 2651
 *Lawrence H. S 2792
 Laybourn, R. L. 2788.
 Leahy F E 2679
 *Leather Makers Process Co., The, 2755
 Lebrun I. 2624.
 Lecog R 2782
 Lee R. C 2702
 Lee W M 2506
 Léger E 2812
 *Legerlots H. 2739
 Legraye M 2670
 Lehmann E 2682
 Lahnartz E 2784
 Lehn A. 2855
 *Lehnert, A. 2800
 Lahrman L 2862
 Leiboff S L 2660
 Leick, J 2790
 Lepuska A 2536
 *Lemaire A. G. 2602
 *Lesgradoff M., 2827.
 *Lesgradoff W., 2827
 Lengyel, B. 2626.
 Leonard, L. T., 2799
 Leostoff, W., 2441
 Lepeshin, V V., 2757
 van Lerbeyge 2635
 Leuche E., 2620
 Leppagott A., 2749
 Levi, A. A., 2724.
 Levi G. R., 2644
 Lewis, I., 2572.
 Levantova, H., 2849
 Levy G., 2754
 Levy J., 2635
 Levy, J., 2765
 *Lavy M., 2817.
 Lewis, B., 2612.
 *Lewis, F S., 2517
 Lichtman, 2765
 *Lichtenberger, T., 2816.
 Liebalky H A., 2637
 Miesche, O., 2606
 *Lilienfeld, L., 2547
 van der Linden A., 2664
 *Linke P G., 2604
 Linton, E P., 2611
 Lintet W 2766
 *Lippmann L. 2513.
 Littenbous E. E 2663
 *Little, A. D., Inc. 2823.
 Littman J B., 2715
 Lloyd T C., 2782
 Lobstein, J E., 2811
 *Locmille, C., 2828
 Loeche, A., 2785.
 Loew, G., 2607.
 *Loewe, S., 2615.
 *Loren, C. F 2690
 Lorigov, N., 2572.
 *Lohmann, C. E., J 2814.
 *Louro J., 2650
 *Lombardi, A. 2639
 van Loon J., 2659
 Lovetz, H. A., 2635
 Loskewitz W., 2675
 Lottemoser A. 2820
 *Low A. & Sons Ltd 2604
 *Low D W 2604
 Lowry T M., 2615 2519
 *Lubricating Products S A 2546
 *Lucas, O D 2861
 Luchs A. 2701
 *Lucian, A. N 2603
 Luck, J M., 2769
 Luckow C 2804
 *Lüdke W., 2823
 Lüppo-Cramer 2632
 Luth, H. C., 2749
 Lühkes, K., 2791
 Lund, E J., 2769
 Lush, R. H., 2781
 *Lutber, F 2831
 Lax H., 2662
 Lyaa, E. V., 2912.
 Lyons, R. E., 2708.
 Maam, O., 2611
 McClelland, C. K., 2795
 McClelland, J F., 2664
 McCombe H., 2705
 McCombe, T H., 2691.
 McCrea, G W., 2615
 McCulloch, A., 2533
 *McCullough, J F K., 2649
 McElvinn, S. M., 2709
 *McGinness, M., 2837
 Mach, P., 2600.
 *Machinery Development Co., 2575
 Macht, D. I., 2756
 *McIntosh, F P., 2690
 *McKee, A.-G., 2603
 McKee, R. H., 2540
 *McKee, E. F., 2651
 MacKenzie K., 2434
 McKibbin, R. J., 2792
 McKinney, P. V., 2699
 McLean, A., 2713.
 McLester J S., 2763
 Macmillan, W G 2705
 MacMurphy F., 2669
 McNear J B 2786
 McNally J G 2623
 *McQueen B. H., 2625
 Madira, V., 2753
 *Madge, E W., 2577
 Magaña H 2744
 Magnus A., 2619
 *Majer A., 2729
 Malhe, A., 2692
 Mamo, M., 2766
 *Major M A., 2823
 Maki T., 2715
 Makray I 2534
 Makunovich, S., 2632
 *Malard P 2661
 *Malet J., 2821
 *Malherbe F J., 2826
 Malquori, G., 2633.
 Mameh, E., 2600
 Mamecke, R., 2753
 Mangan H. A., 2637
 Mangelsdorf, P C., 2762.
 Manjunath B I 2600 2463
 Mannens M. J., 2815
 *Mannesmannrohre Werke 2679 2690
 *Manuch C., 2734
 Marcus A., 2876
 Marascotti, A., 2621 2623
 Margara, R., 2769
 *Margolis J., 2513.
 Maria F., 2553
 Marzocco N., 2620
 van Marle T W J., 2775.
 Marrack, J., 2744
 2767
 Marshall J R., 2649
 Marshall P., 2785

- Martis 2666
 Martin G 2629
 *Martin R H, 2677
 Martin W H 2776
 Marvin G F 2779
 *Mass Ges zur Herstellung Künstlicher überflächen 2632
 Masaki K 2634
 *Maschinen und Fahrzeugfabriken Allet I. Sellgren A I. 2639
 *Maschinenbau A. t. vormals Beck & Henkel 2603
 *Maschinenfabrik Benninger A G 2660
 *Maschinenfabrik M. Gläuser 2647
 *Maschinenfabrik im Jural G. m. b. H. 2602
 *Maschinenfabrik Tilm. Gerber Söhne 2660
 Maschovitz, V. P., 2644
 Mason, C. W., 2635
 Massey, H. & W., 2636
 Masella, P., 2720
 Mast, S. O., 2771
 Masil, R., 2744
 Mathers P. C., 2615
 *Mathieson Alkali Works The 2619
 Mathieu A., 2751
 Mathieu L. L., 2679
 Mathieson, R., 2612
 Matsuki, Y., 2667
 Matsumura, K., 2727
 Matthews H., 2761
 Mathison, A. L., 2664
 Mattill H. A., 2754
 Matzner, M. J., 2747
 Maunther, P., 2607, 2707.
 *Mayer, C. L., 2678
 *Mayer, L., 2677
 *Maynard, I., 2602
 *Mayr, H., 2624
 *Mayr W. M., 2613
 *Mead Pulp & Paper Co., 2651
 *Meier, H., 2605
 Meier, V. W., 2625
 Meier, R., 2752
 *Meier & Welchell, 2679
 Meisner, W., 2626
 Meilen, A. P., 2789
 Mellon, M. O., 2660
 Meisak, J., 2768
 Meiser, W., 2636
 Mendel, L. B., 2762
 Mendelssohn, J., 2633
 Mendelssohn, J., 2778
 Mendlik, F., 2730
 *Menger, A., 2621
 Menne G., 2607
 Menos, J., 2775
 Mercer, P., 2619
 *Merck F. (Hirma), 2606
 *Merckel G., 2639
 *Merrens R., 2642
 Mertle J. H. Jr., 2669
 *Metallger. A. G. 2631, 2678, 2624, 2623, 2676
 *Metals Protection Corp 2619
 Mettkinn R. M., 2707
 Meyer D., 2609
 *Meyer G. E., 2676
 Meyer W., 2609
 *Meyhofer A. P., 2619
 *Milag. Mahlenbau und Industrie A. t., 2781
 Michels W. C., 2640
 Michels J., 2666
 Milauer J., 2676
 *Milke H., 2631
 Miller C. P., 2661
 Miller D. G., 2630
 Müller, V., 2674
 Müllig A., 2777
 Mills W. H., 2696
 Milner, H. W., 2625
 Milner R. T., 2642
 Milskil, A. B., 2673
 Minchin S. T., 2610
 *Minimax A. G., 2623
 *Minokki G., 2617
 Mints I. H., 2672
 Mitchell C. A., 2661
 Mitchell P. H., 2677
 Mitchell T., 2601
 *Mitchum C. P., 2751
 *Mitteldeutsche Stahlwerke A. G., 2602, 2624
 Miyajima H., 2719
 Miyata M., 2651
 *Mohr, A., 2679
 Mohrnoth V. O., 2710
 Moll, P., 2784
 Molloy, J. J., 2662
 Montgomery, J. P., 2606
 *Montoussé, A., 2606
 *Moore W., 2601
 *Moorman A. R., 2644
 *Moongraher, P., 2649
 van Moppe, R. M., 2670
 *Mordey, W. M., 2672
 Morgan, G. T., 2704, 2636
 Morgen H. H., 2662
 Morgan, W. L., 2626
 Morgan, W. T. J., 2767.
 *Morrow, J. E., 2619
 *Mortrud, E., 2647.
 *Morton, R. & Co., Ltd., 2604
 Mory, A. V. H., 2782
 *Moschel, W., 2614
 Mossop, M. C., 2601
 *Motor Improvement, Inc., 2602
 Mouncey J., 2769
 Mosner, N. H., 2639
 *Muehke J., 2702
 Müller 2794
 Müller A., 2604
 Müller C. A., 2677
 Müller D., 2745
 *Müller H., 2607, 2623
 *Mueffler J., 2657
 Müller J. A., 2604
 *Müller O. A., 2650, 2675
 *Müller P. H., 2601
 *Müller R. P., 2640
 Müller R. H., 2624
 *Müller W. I., 2619
 Münch A. W. P., 2781
 Muehlin G. R., 2631, 2632
 *Muller H. J., 2670
 Muller G. I., 2767
 *Muller U., 2647
 Muller C. H., 2616, 2654, 2660
 *Mullinow P. J. & Co., 2634
 Munch J. C., 2617
 Muesout H., 2650
 Murphy D. W., 2629
 *Murphy P. A., 2677, 2678
 Murphy O. M., 2631
 Murphy T. F., 2653
 Musgrave, G. W., 2791
 *Myt. O., 2650
 *N. V. de Bataafche Petroleum Maatschappij 2731, 2732, 2766, 2637
 *N. V. Carbo Union Industrie Maatschappij 2639
 *N. V. Chemische Fabriek L. Van Der Grinten 2637
 *N. V. Maatschappij tot Beheer en Exploitsie van Dettin oleu 2626
 *N. V. Maatschappij tot Exploitsie van Hievinlingen 2626
 *N. V. 't Hiltje' Glaz. lampenfabriek 2669
 *N. V. Follra F'n Oven Louw Maatschappij 2638
 *Naamloze Vennootschap W. A. Scholten & Chemische Fabrieken, 2621
 Naehring, H., 2671
 Naehagen, H., 2694
 Nagasaki, N., 2631
 Nakamura, K., 2606
 Nanji H. R., 2695
 Naeedham, M. N. I., 2657.
 Naayana N., 2664
 Narayanamurti, D., 2741
 Nairne, B. A., 2771
 *Nathansohn A., 2660
 Naylor H., 2796
 *National Carbon Co., Inc., 2614
 *National Processes, Ltd., 2617, 2614
 *Nougatuck Chemical Co. The 2740
 Nuroy A., 2663
 *Nowinsky P., 2657
 Nealey J. H., 2616
 Nederselk H. V., 2713
 Neer, J. B., 2677
 Nehndor, V. A., 2744
 Neil R. H., 2647
 Nelson C. M., 2761
 Nelson V. I., 2762
 Nende A., 2794
 Nemes T., 2676
 Nemetz I., 2770
 Neomeyano A. N., 2702
 *Netsch G., 2627
 Neumann H., 2634
 Nettschloss R. M., 2745
 *Newton P. D., 2676
 Newnan R. H., 2782
 *New Jersey Zinc Co., The, 2678
 *Neynaber, K., 2636
 *Nienlandt J. L. P., 2637
 Nierenstein, M., 2719
 *Niessé H. A., 2602
 van Nieuwenburg C. J., 2657
 Nieuwen, J. A., 2641, 2692
 Nieuwenhousen H., 2640
 *Nikolai, P., 2627
 Nilsson, R., 2744
 Nishimura, H., 2745
 Nishikawa, K., 2667
 Noble, I. T., 2774
 *Nock, J. A., Jr., 2660
 Nogata G., 2614, 2655
 *Nohe, W., 2785
 *Nordstrom, B. J., 2605
 Nordt, H., 2607
 North, G. W., 2779
 North, R. V., 2741
 North, R. G. W., 2641
 Nouvel, O., 2664
 Nozak I., 2609
 Nucconi, R., 2794
 2796, 2607.
 *Nume, A. de Saint Edme Horne Honet, 2603
 *Nuttall, T. H., 2651
 Nyli, W., 2612
 *Oberbach, J., 2611
 Oberg P., 2705
 Oht, P., 2600
 Oka, R., 2641
 *Okida Smelterwerk Ak. Gesellschaft, 2602.

- Odland T E 2400
 Oelrich C 2734
 Oelverberg W 2339
 Ogibetta H 2447
 von Ohlen F W 2757
 Ohlhauser H 2781
 Okada A 2680
 Okamura T 2778
 Oku M 2788
 Oleovich H S 2758
 Oldright G L 2674
 Oldroyd C E 2460
 Olus J F 2708 2722
 de Oliveira R P 2670
 Olson G 2673
 Ono M 2719
 Onslow M W 2758
 Ormandy W R 2441
 Orntun L S 2606
 Orochko D I 2725
 van Os D 2869
 Osgood E E 2781
 Osgood F D 2833
 Osterberg A E 2753
 Osterhout W J V 2736
 *Osterreichische Chemische Werke, G m b H 2461
 *Osterstrom, R C 2848
 Ostrupovich, A 2739
 Otto, C 2747
 Otto, C 2839
 Otto, C & Co, G m b H 2821
 von Overstraten C 2861
 Owen W D 2782
 Ozaki J 2784
 Packer, J 2698
 Pades, M 2643
 *Pahl, A 2802
 Pannwitz E 2830
 Pap L 2772 2773
 Parker, J 2848
 Parkes A S 2768
 Parlow A 2873
 Parry, E J 2813
 *Parry H J 2803
 Parsons, A L 2667 2668
 Partridge E F 2682
 Partridge H M 2628
 Paschke V 2644
 Pascoe E H 2668 2818
 *Passerman M 2603
 Pastor I 2776 2777
 *Patentaktiebolaget Gröndal Raima 2677
 *Patent Treuband. Ges für Elektrische Glühlampen 2651
 Patnode, W I 2639
 Patry M 2852
 *Pattillo, D K 2851
 Pätzler, J 2753
 *Peebles D D 2851
 *Peebles, T A 2604
 Peers, H 2854
 *Penherton R T 2792
 Penney, W G 2643
 Penrose, M 2743
 Pensler W H A 2833
 *Perda R 2827
 Perman H F 2734
 Perotti, R 2796
 Perrott, G St J 2612 2652
 Perry E S 2787
 Price B 2613
 *Peter L H 2643
 Peters K 2692
 Peters, R A 2760
 Peterson W R 2645
 *Petroleum Conversion Corp 2844
 Petter, F W 2401
 Petzall, R 2782
 Pfuhl E 2783
 Philbeck, F A 2633
 *Philipp, O (Farma) 2603
 Phillips, F C 2672
 Phillips W M 2647
 Philipp, T E 2600
 *Physical Chemistry Research Co 2637
 Piatti L 2782
 Pichot J 2612
 Picon M 2624
 Pieper, R 2733
 Pierre W H 2794
 Pieters H A 2813
 Pilgrus, G T 2696
 Pinusson L 2761
 *Pintsch J A-G 2602
 Piotrowski, W J 2841
 *Pipkin S 2846
 *Pismarev S 2832
 *Piza, D 2604
 Plattmann C R 2828 2830
 Piodelli I 2802
 Pinea L N 2833
 Pisch J 2747
 Plotnikov V A 2644 2645
 Pöcher W 2763
 *Poulet, Ltd 2803
 Poboň F 2675
 Podbielniak W J 2613
 Polnard E 2670
 Polomovsky M 2749
 *Polymer G A G 2833
 Pomeranz H 2869
 Ponder E 2780
 Pomeroy N 2778
 Pontoppidan C 2631
 Popov A M 2707
 Popper H 2747
 *Porritt U & Co A G 2602
 *Poscotta A 2679
 Potter A 2789
 Powell, G 2706
 Powers, H D 2765
 Prager, H Jr 2471
 Pratt O R 2780
 Praxmair, F H 2649
 Press, A 2612
 Price F H 2719
 Price W A 2760
 Prudeaux, E B R 2819
 *Prilling, K 2406
 Price F H 2601
 Probst, A 2707
 Probst, A 2743
 Prodan, L 2661
 Proft, E 2701
 Promieser N 2646
 Promer, M 2409
 Promer, C W 2409
 Pütter, A 2766
 Pugsley, L L 2792
 Pufke, S L 2622
 de Putnoky L 2826
 Putnoky L 2708
 Putzky, P 2627
 Pyball, H 2432
 *Quartz et Salze 2428
 Quastel J H 2741
 Quader A 2693
 Rabas, A 2635
 Rabale H 2462
 Robinson A I 2620
 *Rabot, L V 2433
 *Rademacher, W 2792
 *Radó, L 2634
 *Radt, C 2614
 Radt C 2728
 *Ragan F H 2826
 Rahn O 2745
 Ralford L C 2701
 *Rath J 2831
 Rakusa M A 2695
 Rao A J 2407
 Rawen G 2768
 Rawney A A 2779
 Rawnsley G A 2778
 Ramsperger H C 2630
 *Randall F. (Luton) Ltd 2409
 Randon I (Aime) 2762
 Rankine A O 2612
 Rao R 2401
 Rao, S V 2863
 Rapoport I B 2433
 *Rambach W 2818
 Ressel B 2785
 *Rathburg H 2853
 Raudel 2630
 Rauler Z 2605
 Ray P R 2635 2638
 Read, W T 2606
 Reader V 2769
 Redman L V 2782
 Reed C I 2761
 Reed H R 2790
 Reed H S 2757
 *Reed, M C 2876
 Reed R F 2632
 *Reeson, F A 2850
 Rege, R D 2798
 de Regny F V 2771
 Reich, G T 2816
 Reickart, H L 2777
 *Reisman F R 2823
 Reiner, M 2613
 *Reiss A 2783
 Reistätter J 2649
 Remington R E 2664
 Renaudie 2682
 *Remmeret, L 2640
 Reunk J B 2783
 Rettger, L F 2798
 *Revos, W 2634
 Révy, D 2622
 *Rewald B 2777 2741 2743
 Rym, F R 2407
 *Reymersholms Gamle Industri Aktiebolag 2678
 Rheinbold, H 2691
 *Reinhold & Co, Vermehrte Kieselguhr und Kalksteine, Ges 2423 2431
 *Rheinische Kampfer Fabrik G m b H 2734 2740
 *Rheinisch Westfälische Kalkwerke 2820
 Richards, F J 2789
 Richardson, A S 2605
 Riddell, W C 2830
 Riddle O 2782
 Ridenour G M 2789
 *Riebeck'sche, A. Moutanwerke A G 2845
 *Riedel J D & de Haen A-G 2734 2736 2740
 Rietze O 2876
 *Rigault L 2823
 Rigoni M 2744
 *Riley R H J 2850
 *Riley, T N 2787
 *Rimington C 2746
 Ring L W 2813
 *Ritschel O 2792
 *Ritschl R 2827
 Rivie C W 2782
 Robertson G R 2816
 Robae, H H 2674
 Robinson, G W 2792
 Robinson M E 2747
 *Robinson P 2604
 *Robson S 2817
 Roche J 2769
 Rodé E Y 2628 2646
 Rodebusch, W H 2836
 Rodewux A 2832
 Roe, L I 2718
 *Röhm, O 2878
 *Röhm & Haas Co (Soc Anon) 2803 2824 2861 2875
 *Röhrig H 2779
 Roelen W 2433
 *Romer R 2604
 Rogal, F 2792

- Rogers L. A. 2804
 Rogoff J. 2761
 *Ale Robden C. 2821
 Rojahn C. A. 2808
 Rolla L. 2608
 Roller P. S. 2529
 Roman W. 2751
 Romwaller A. 2833
 Rona I. 2637
 Ronve H. 2752
 Roscoe M. H. 2760
 Rose I. H. 2696
 Rose W. H. 2762
 Rosebury T. 2760
 2761
 Rosenberg H. A. 2761
 *Rosenberg F. 2614
 Rosenblum N. 2637
 Rosene H. F. 2769
 Rosenmund K. W. 2682
 *Rosenstein O. 2626
 Rosenthaler L. 2604
 *Roses - Galceran P. 2831
 Ross, J. 2659
 *Ross, P. 2602
 Ross, G. 2621, 2623
 *Rothenbrunner Seifen- und Ölfabrik Schmieg & Schieb G m b H. 2670
 *Rothra, W. S. 2740
 *Roucka, I. 2604
 Roulala, O. 2663
 Ruess A. W. 2761
 Rowe F. M. 2684
 Rowell, H. W. 2617
 Rosas M. L. 2671
 Roy, S. (Miss). 2622
 *Royal Baking Powder Co. 2604
 Royanski, D. 2608
 Rzyer, L. 2616
 Rozanov, S. N. 2664
 Rozroff, A. D. 2611
 *Rudolfs, W. 2792
 Ruff O. 2633
 Rugg's P. 2722
 Rule, H. G. 2715
 *Rumbarger, B. W. 2639
 Rumbold W. G. 2674
 Rummel K. 2601
 *Rummerl K. 2604
 *Runge, W. 2637
 *Rupp F. 2603
 Ryan, W. A. 2790
 *Rymal, C. H. 2792
 Saam, H. G., Jr. 2750
 *Sachs, A. P. 2644
 *Sadkay, V. S. 2700
 *Sadtler, H. D. 2823
 *Sadtler, H. S. 2663
 Saha, H. 2653
 *Saudershan, N. 2678
 St. Josesco, 2755
 St. Reiner, 2600
 Sakuma, H. 2667
 *Salles, P. M. R. 2649
 Salmon, W. D. 2763
 Sambamurti G. 2637
 Sanders, K. 2601
 2780
 Sanders, P. R. 2788
 Sandford I. 2762
 Sandoonni C. 2668
 Sandoz S. 2611
 de Sankor Z. 2772
 Sand G. 2653
 *Sandig, R. B. 2646
 Sanna G. 2720
 Santos, A. C. 2607
 Saran W. 2677
 Sasaki, S. 2743
 Saslow G. 2750
 *Saxe W. 2602
 *Shtava, J. 2604
 Sato K. 2754
 Sauer E. 2674
 Sauerbrey W. 2673
 Saunders B. C. 2696
 *Savargos J. 2539
 Savard J. 2702
 Sayers, R. R. 2783
 Seaborough H. A. 2705
 Seaboard, G. 2631
 *Schack A. 2604
 Schaefer, C. 2635
 Schaefer, J. 2629
 Schaller, W. 2693
 Scharrer, K. 2763
 Scheller, F. B. C. 2632
 Scheifele, B. 2664
 *Schepk, P. 2603
 *Schering - Kahlbaum A. G. 2736, 2740, 2603, 2614
 Schiebl, K. 2672
 Schiebold, E. 2643
 *Schlaret, K. 2633
 Schleede, A. 2643
 Schlichter, S. 2675
 *Schless, S. 2632
 Schlick, W. 2664
 Schlüter H. 2617
 Schluett, H. 2637
 Schmalz, P. J. R. 2764
 Schmatolla O. 2609
 *Schmies, H. 2570
 Schmidt, A. 2634
 *Schmidt, B. 2782
 Schmidt, E. A. W. 2637
 *Schmidt, K. 2659
 Schmidt, L. 2672
 *Schmidt, W. 2645
 *Schmidt - Aerosol G m b H für Feuer- ausreglung, 2604
 *Schmidt'sche Heilm- dampf Ges. m b H. 2604
 *Schmieder, J. M. 2645, 2651
 Schmitt, W. 2765
 Schmuts, A. 2742
 Schneider, E. 2643
 Schneiderböhne, H. 2670
 *Schnurpfel, F. 2802
 Schöberl A. 2633
 *Schosler, A. 2677
 Schorn, M. 2673
 Schonebeck v. 2677
 Schönfelt H. 2665, 2669
 *Schöncke W. 2660
 Scholten W. P. 2631
 Schoonover, J. W. 2747
 Schost N. 2661
 2665
 *Schopper A. 2603
 Schreiber L. 2663
 Schreimakers, P. A. H. 2623
 Schreier H. H. 2743
 *Schroder H. H. 2516
 *Schroeder M. 2617
 Schroeder R. A. 2667
 *Schroinger E. 2636
 *Schubert P. 2631
 Schuette C. N. 2674
 Schuthof, L. 2743
 *Schulte, A. 2741
 Schulte G. 2633
 Schula, P. 2672
 *Schulze, E. 2514
 Schumacher, H. - J. 2630
 *Schumann, L. 2631
 *Schurer, M. B. 2544
 Schwab, R. 2742
 Schwaibold, J. 2772
 Schwalbe, C. G. 2630
 *Schwarz G. 2632
 Schwenner, H. 2615
 *Schwarz, C. 2679
 Scott, D. A. 2741
 Scott, E. B. 2674
 *Scott, T. R. 2747
 *Scottish Dyes, Ltd. 2735, 2738, 2634, 2639
 Seeger, B. 2633
 Seel H. 2759
 Seehrer, R. 2647
 Seeger, E. 2641
 Seitz, E. 2743
 Seitz W. 2652
 *Schumann, A. 2745
 Sell, W. 2599
 Semichon, L. 2665
 *Semson, W. L. 2576
 Sen, H. 2634
 *Senfter, H. 2631
 *Senfter, G. 2782
 Sentense, 2790
 Sera, S. 2710
 Serkin, E. S. 2616
 Setlik, B. 2647
 Shaglow, A. Y. 2700
 *Shakespeare, W. M. 2632
 Shakhov, A. P. 2633
 Shank, M. 2647
 Shannon, R. C. 2791
 Shapley, H. 2643
 Sharp, T. M. 2730
 *Sharples Specialty Co., The. 2645
 *Shaw, C. 2636
 Shaw, E. J. 2600
 Shaw, H. 2718
 Shaw, R. 2623
 *Shchukina M. N. 2775
 Sheets G. 2760
 Shenon F. J. 2669
 Shenstone A. G. 2640
 Shenston H. C. 2790
 Sheppard, R. B. 2623
 Sherman, A. 2641
 *Sherman A. E. 2602
 Sherwood R. C. 2772
 Shibuya C. 2710
 Shlatre V. P. 2607
 Shoemaker, B. H. 2644
 Shokhet, I. I. 2572
 Shorvitz, P. P. 2640
 Shiner R. L. 2636
 Shupe, L. R. 2612
 Shure A. P. 2600
 *Sicard, A. 2624
 *Sicard, J. L. 2624
 Siebert G. 2667
 *Siemens - Elektro Osmose Ges. 2603
 *Siemens & Halske A.-G. 2631, 2619, 2624
 *Siemens - Mann- werke A.-G. für Kühle fabriken 2650, 2651
 *Siemens - Schuck wertwerke A.-G. 2603, 2631, 2661, 2756, 2650
 *Siemens - Vefa Ges. für Medizinische Technik, 2603
 *Sokolowski, S. 2764
 Solovov, V. 2624
 *Solena Verrie chem- ische Fabriken, 2614
 *Solter & Rolien kirchen G m b H. 2631
 da Silva Araujo, C. 2606
 *Sammon, H. 2604
 *Samon A.-G. 2670
 *Samon Brothers (Engineers), Ltd. 2570
 *Samon L. J. 2570
 Samonnet, H. 2763
 *Sampson, E. 2577
 *Sand, O. 2650
 Sanelnikov, K. 2608
 *Sanger, P. 2679
 Sangh, D. 2711
 Sangh, M. L. 2711
 *Sangster, J. A. 2649
 Sipyagin, A. S. 2616
 Skaar, J. R. 2576
 *Skappel, H. 2677
 Skilling, W. J. 2633
 Skinnor, D. G. 2634
 Slater, C. S. 2793
 Sifer, E. H. 2771
 Slotta K. H. 2613
 Smekal, A. 2614
 *Smeltverck L. A. S. 2619

- Amenda V 2535
 Amies S 2524
 Smith A L 2759
 Smith C O 2622
 Smith D P 2642
 Smith H H 2662
 Smith I L 2760
 Smith R R 2667
 Smith I C, 2744,
 2767
 Smith G 2674
 Smith G F 2654
 Smith H G 2623
 Smith H O 2635
 Smith J F D 2613
 Smith J H C 2623
 Smith J N 2649
 2651 2787
 Smith J W 2663
 Smith P W 2624
 Smith, R C 2477
 Smith S S, 2541
 Smith S W, 2670
 Smith W R, 2666
 Smithell C J, 2672
 Smith, A, 2613
 Smoleysk E 2743
 Smorodintsev, I A,
 2745
 Smotter, M C 2678
 Snell F D, 2416
 2429
 Snyder, N H, 2433
 Snod R, 2744
 Soc anon des dis-
 tilleries des Dents
 2406
 Soc anon des établis-
 sements Armeal
 Braumstien, 2532
 Soc anon pour lind
 chim A Hle 2766,
 2815 2921, 2931
 2945 2956, 2957,
 2958 2960 2965
 Soc anon des manu-
 factures des glaces
 et produits chimi-
 ques de St Gihann
 Chauny et Cuy,
 2926 2927, 2928
 Soc anon m Neel &
 Cie, 2735
 Soc anon des procé-
 dés R Audubert
 2461
 Soc electro - metel
 lorsque de Saint
 Héron 2631
 Soc d'études chim
 pour lind, 2903
 2819
 Soc d'études scien-
 tifiques & d'entre-
 prises industrielles
 2785 2818
 Soc pour la fabri-
 cation de la soie Rho-
 diensis 2949 2962
 Soc générale des fours
 à coke systèmes Le-
 cocq 2817
 Soc Gustav Wippen-
 mann Maschinen
 fabrik Stahlwerk
 und Eisengieserei
 G m b H 2439
 Soc Hl Chabal et
 Cie, 2602
 Soc industrielle des
 dérivés du soufre
 2873
 Soc des matières
 colorantes et prod-
 uits chimiques de
 St Denis 2653
 Soc des phosphates
 Tuscienne et des en-
 grais et produits
 chim, 2317
 Soc des produits
 chim de Clamecy,
 2138
 Soc Radierwerk E.
 Schrach 2603
 Soc de recherches et
 d'exploitations pé-
 trolières, 2604
 Soc Seedorf - Har-
 tord Georges 2427
 Soc des mines chimi-
 ques Rhône Pou-
 lenc 2427
 Soc Villedu Fabre,
 2403
 Sokolova M N, 2617
 Solinger, F, 2700
 Solon K, 2472
 Solz, L, 2709
 Sommer H H, 2775
 Sommer, J, 2446
 Sone, C, 2727
 Soni C L, 2636
 Soonswaile, N F, 2607,
 2608
 Sornet R, 2755
 Spence, H S, 2634
 Spencer, C M, 2876
 Spencer G C, 2659
 Spencer H M, 2444
 Spencer W D, 2631
 Spensler O, 2570
 2572
 Spillers, Ltd, 2633
 2445
 Spielman M A, 2686
 Spoerri & Co., 2423
 Sponer, H, 2634
 Spoon I W, 2543
 Spradlin, M C, 2750
 Sprenger G 2630
 Sprouton F 2566
 Srinasa M M 2703
 Sreenivasaya, M
 2468
 Staacke C W 2476
 Staib K 2631
 Stalger 2404
 Stalkorski K M
 2629
 Stammreich H
 2785
 Stampe G 2744
 Standard Telephone
 & Cables, Ltd, 2631
 2757
 Stanek V, 2601
 Stansel N R 2644
 Stepp P, 2820
 Stark C, 2847
 Stark, W E, 2605
 Starke A, 2464
 Start, I. Jr, 2750
 Steay, T D, 2679
 Steckhorn Kunst-
 seide A G, 2844
 Steele, C C, 2733
 Stein E, 2404
 Stein, O, 2612
 Steinberg F, 2659
 Steinhilber, A H, 2761
 Steimetz, H, 2616
 Stempel O H, Jr,
 2708
 Stengel, L A, 2736
 Scropoe A., 2420
 Stephan, K, 2710
 Stephenson J, 2451
 Stephen, H, 2680,
 2710
 Stephens, D W, 2692
 Steppan O, 2510
 Stern R, 2644 2873
 Stern T E, 2608
 Steithacher, A, 2453
 Stetter, O, 2637
 Stettiner Chamotte-
 Fabrik A G vor-
 malis Dührer, 2603
 2625
 Steudel, H, 2743
 Stevens, H P, 2676
 Stewart K. P., 2789
 Stewart, F D, 2689
 Stewart, F H, 2793
 Selasny, H, 2414
 Saliger, G, 2631
 Selles, J C, 2734
 Selles W, 2623
 Sellmwell, C. W, 2606
 Sells, J, 2774
 Seck A, 2662
 Seckhausen, A H,
 2674
 Seewater F 2822
 Seid, R, 2721
 Seitzberg H, 2730
 Seiser F H, 2651
 Seiser G R, 2618
 Seiser, F D, 2702
 Seisquedam R
 2431
 Seranski J N, 2619
 Seucker O C
 (Furnal), 2445
 Seubling R M 2469
 Strass J, 2599
 Stucky, C J 2762
 Staff, F 2651
 Stamper R 2613
 Stucky R 2631
 Söhr J 2731
 Suffer G, G 2609
 Sugar Beet & Crop
 Devere, Ltd 2373
 Sugata, M 2522
 Sullivan F W, 2845
 Sullivan J D, 2673
 Sullivan, M X, 2745
 2749
 Sulphide Corp, 2677
 Sumiki Y 2734
 Sundin, N, 2667
 Superfine Chemicals,
 Ltd, 2821
 Surányi, J, 2799,
 2821, B, 2758
 Suter, C. M, 2703,
 Suter, H, 2549
 Suthers, A J, 2715
 Suthers, W D, 2734
 Sutton, T C, 2399
 Sutton, T S, 2778
 Svatek E, 2750
 Svedberg T, 2746
 Sesacow K., 2644
 Sverrev, V, 2310
 Swan Hunter & Wg
 ham Richardson
 Ltd, 2846
 Swart, E L, 2612
 Swartout, H D, 2750
 Sweeney, O R., 2822
 Swietlawski, W,
 2607
 Syndicat des laitières
 et acorles 2678
 Synges, E. H, 2619
 Szaffke T, 2833
 Szanyi, I., 2775 2777
 Szayna A, 2340
 Szeghari, A, 2876
 Szelen, A, 2401
 Takai, S, 2719
 Taketomi, N, 2804
 Tamachika, M, 2850
 Tamchyna, J V, 2601
 Tan J P, 2671
 Tancet, G, 2763
 Taylor, H, 2763
 Taylor, H S, 2641
 Taylor Wharton Iron
 & Steel Co., 2660
 Technicolor Motion
 Picture Corp, 2653
 Teller, S. V, 2758
 Teltsch & Co, E.,
 2472
 Tummkova, T. I.,
 2712
 Teorell T, 2620
 Terau, A, 2772
 Terfayl, S, 2402
 Terwille, J, 2617
 Tesch & Stabenow
 Internationale Ges
 für Schädigungsbe-
 ampfung m b H.,
 2403
 Teufel, O, 2673
 Teunissen, H P, 2715
 Thecker R A, 2701
 Thate H, 2726
 Thaus, A, 2458
 Theiss, L, 2643
 Ther, R, 2460
 Thermal Industriest &
 Chemical (T. I. C.)
 Research Co., Ltd.,
 2830
 Thieme, P A, 2600
 Thimano K V., 2623
 Thomas & Co., 2785
 Thomann, J, 2813
 Thomas, A. A, 2605
 Thomas & Hochwalt

- Laboratoire Inc., 2467
 *Thomas, J., 2856, 2859
 Thomas, J. O., 2769
 Thomson, A. S. T., 2812
 Thomson B., 2667
 *Thomson G. M., 2664
 *Thomson, R. P., 2856
 *Thornton J. P., 2654
 Thornton, N. C., 2766
 Thorpe, J. P., 2693
 *Thorsell C. T., 2820
 *Thyssen'sche Gas- u. Wasserwerke, G. m. b. H., 2603
 *Tidmore, A. H., 2850
 Tiesens, O. J., 2706
 Tikhomolov, P. A., 2712
 *Tilley, R. D., 2681
 Timm, W. B., 2663
 Timmerman, W. A., 2767
 Tjaden, 2778
 Todd, P. W., 2763
 Tolch, N. A., 2852
 Tomita, M., 2731
 Tompa, B., 2631
 Toporescu, P., 2624
 Toussaint, A., 2686
 Towne, C. C., 2812
 Townsend, J. B., 2639
 Traik, P. D., 2671
 Traub, I., 2616
 Travis, A. J., 2700
 Treloar, A. F., 2774
 Tremals H. E., 2664
 Triaks, W., 2601
 *Trinquart, L. G., 2618
 Trivelli, A. P. H., 2652
 *Trochungs - Verschmelzung- und Ver- gassungs-G. m. b. H., 2837
 Troque, C., 2623
 Trotman, P. E., 2751
 *Trensdale, R., 2877
 Trumper, M., 2783
 *Truder, R. B., 2867
 *Trust, A., 2654
 Tsao, S. L., 2670
 Tschirch, R., 2665
 *Tucker, R. T., 2815
 Tocker, S. H., 2702
 Tuminskattl, M. C., 2807
 *Turek, P., 2602
 *Tustian, O. R., 2831
 van der Tuuk, J. H., 2632
 *Twiss, D. P., 2877, 2878
 Tyler, P., 2609
 Tyler, F. M., 2816
 *Tyrer, D., 2821
 Tyrrell, G. W., 2672
 Ubbelodhe, L., 2868
 Ugoln, O. B., 2614
 Uhl, G., 2462
 *Ulmann G., 2460
 Umstatter, H., 2618
 *Union chimique Belge
 Snc anon., 2440
 *United Water Soften-
 ers Ltd., 2611, 2792
 Urbach P., 2619, 2647
 *U. S. Industrial Alco-
 hol Co., 2847
 Ure la V. D., 2784
 Vabiyannathan V. I., 2691
 Vallerster R. B., 2634
 Van Der Meer, C. R., 2692
 Van derburg W. D., 2672
 Vandone, I., 2812
 Vansell, I. H., 2778
 Varga, J., 2812
 Varga, L., 2797
 (Varga J.) 2780
 Vas K., 2713
 *Vaughan C., 2679
 Vaughn, J. M., 2767
 Veibel, S., 2743
 Veik, S., 2632
 von Velmann N., 2619
 Velasco, M., 2693, 2614
 Velick, J., 2644
 Vencov, S., 2334
 Venkateswaran, B., 2641
 Vercelli, C., 2671
 *Verrein für chem. Ind.
 A.-G., 2818
 *Vereinigte Alumin-
 ium Werke A. G., 2681
 *Vereinigte Glasa-
 stoff Fabriken A. G., 2818
 *Vereingte Stahlwerke
 A.-G., 2691, 2677, 2679, 2680
 Vermeulen D., 2606
 Verotte, P., 2633
 Verona O., 2797
 *Vickers, Ltd., 2861
 Vickery, H. B., 2809
 *Vidal, R., 2678
 Vito, M., 2643
 Vite, P., 2664, 2768
 Voigt, J., 2623
 Voigt, O., 2600
 Volkowich, B. I., 2799
 Vondráček, R., 2664
 de Voogd, J. G., 2664
 Voopl, N. H., J. M.
 (Mile), 2632
 van Vorst, F. T., 2816
 *Vorländer, D., 2820
 *Wacker A. Ges. für
 elektrochem. Ind. G.
 m. b. H., 2734, 2647
 Waddington C., 2639
 Waever, B., 2658
 Wagner, F. C., 2606
 Wagner, O. H., 2642
 Wnquer, W., 2864
 Walden, F., 2626
 Walkee D. J., 2753
 Walker, P. F., 2803
 Walker, R. H., 2799
 Walker, S., 2785
 Walker, T. K., 2718
 Walker T. I., 2647, 2692, 2674
 Wallace J. H. Jr., 2659
 Wallace R. H., 2784
 Walsh P. C., 2816
 Walter G., 2808
 Walter J., 2816
 Walton P., 2704
 Walton, S. T., 2768
 Wampler R. W., 2617
 *Wappler P. H., 2601
 Ward J. C., 2813
 *Waring H., 2692
 *Wark N., 2679
 Warner T., 2670
 *Wartenberg H., 2633
 Wartsman P. B., 2672
 Watanabe M., 2693
 Waterman H. S., 2617
 Waterman H. I., 2608
 Waters W. A., 2632
 *Watson C. B., 2644
 Watson, K. M., 2613
 Waybitt H. G., 2606
 Weber, H. H., 2666
 *Weel J. M., 2681
 Weegmann P., 2641
 Well, L., 2812
 Weinfield, A., 2670
 Weiss, A., 2783
 *Weisheit, G., 2789
 Weiss J. J., 2631
 Welch, K. N., 2691, 2693
 Werner, H., 2647
 Wertheimer, E., 2742
 Weson L. G., 2747
 West O. D., 2611
 Westenberg, L., 2714
 *Westinghouse Lamp
 Co., 2693
 *Weston Paper and
 Manufacturing Co.,
 The, 2851
 Weston R., 2791
 Wettstein A., 2731
 Wever, P., 2699
 Whang, B. H., 2621
 Wheaton R. S., 2777
 Wheeler, T., 2762
 Whitacre, P. M., 2689
 Whitaker, R., 2778
 White P., 2606
 White, P., 2874
 Whittier, P. O., 2804
 Whitton, J. P., 2684, 2714, 2721, 2725, 2730
 Wichers, E., 2659
 Wiek, P. G., 2643
 *Wie Wierwinsky, A., 2893
 Wiebe, R., 2607
 *Wiedemann, R., 2802
 Wieringa, K. T., 2746
 van der Wielen, P., 2898
 Wierma P. C., 2699
 Wigginton, R., 2832
 Wigner, R., 2643
 Wilczewski, P., 2809
 *Willich J. A., 2692
 Wilkin, P. J., 2631
 Wilkinson, P. D., 2762
 Williams A., 2670
 Williams J. W., 2624
 Willock H. H. T., 2723
 Wilson C., 2790
 Wilson P. J., 2701
 Wilson S. P., 2664
 Wimer D. C., 2796
 Winick, H., 2616
 Winkler J., 2810
 Winkler L. W., 2633
 Winkler W., 2781, 2814
 Winograd, A., 2644, 2694
 *Winter R. M., 2695
 *Winterfeld K., 2736
 Winters P., 2796
 *Winterstein A. H., 2850
 Witt P., 2718
 *Wittich H., 2803
 *Wittmeier H., 2601
 Wohine, R., 2743
 Wolf, H., 2693
 Wolfenden, J. H., 2623
 Wolf, H., 2853, 2844
 *Wolfner, J., 2814
 *Wollsch, K., 2649
 Wollin K., 2741
 *Wolner, P., 2827
 *Wolner, R., 2827
 Wowl, L. J., 2619
 *Woodall Durham
 (1920), Ltd., 2663, 2839, 2810
 Woodson, J. C., 2644
 Wood, P., 2613
 Woolcock, J. W., 2626
 Work, H. K., 2616
 Woods, J., 2606
 *Woudhuysen, J., 2846
 *Worwick, O., 2747
 Wratschko, P., 2621
 *Wrighton, W. J., 2681
 *Wundermann Objektiv-
 optische Chemische
 Fabrik Schlenker
 "Wasschlimp" 2821
 *Wülling, J. A.,
 (Firma), 2814
 Wwart, J., 2614
 Yamada M., 2688
 Yamamoto, K., 2849
 Yamauchi, Y., 2661
 Vancey, H. P., 2831, 2834
 Vant, W. P., 2783
 *Young H. J., 2840
 Zaltschek, A., 2789
 Zalkind, Y. S., 2719, 2717, 2714
 Zaitman, W. H., 2713
 Zechner, L., 2809

CHEMICAL ABSTRACTS

Editor: E. J. Crane, Ohio State University, Columbus, Ohio.
Associate Editors: Elmar Hockatt, Helen Gama and Leonard T. Capell
Assistant Editors

M. S. Anderson	D. J. Demarest	P. B. Hawk	L. A. Olney	J. J. Skinner
W. L. Badger	Alden H. Emery	P. R. Howe	A. M. Patterson	W. R. Stemen
G. E. Barton	W. O. Emery	C. H. Karr	T. G. Phillips	C. O. Storm
Edward Bertow	W. F. Faragher	H. A. Lepper	A. N. Richards	E. M. Symmes
S. R. Benedict	A. C. Fieldner	R. K. Marshall, Jr	Richard Rimbaech	P. P. Underbill
F. C. Blank	C. G. Fink	A. R. Middleton	Allen Rogers	H. G. Wells
F. L. Browne	C. N. Frey	Harlan S. Miner	C. A. Rouiller	C. J. West
C. E. Curran	H. W. Gillett	Chas. B. Murray	A. H. Sabin	H. T. Wherry
J. K. Dale	R. A. Gortner	C. H. Munroe	J. P. Schairer	R. F. Wightman
C. C. Davis	W. T. Hall	W. A. Nayer, Jr	R. Scherubel	J. C. Witt

CONTENTS

Patents	1	13. The Fermentation Industries ..	3140
Author Index	1	17. Pharmaceutical Chemistry ..	3144
1. Apparatus and Plant Equipment ..	3374	18. Acids, Alkalies, Salts and Sub-	3151
2. General and Physical Chemistry ..	3434	dries	
3. Subatomic Phenomena and Radio-		19. Glass, Clay Products, Refractories	3143
chemistry	3410	and Enamelled Metals	
4. Electrochemistry	3123	20. Cement and Other Building Ma-	3144
5. Photography	3123	terials	
6. Inorganic Chemistry	3120	21. Fuels, Gas, Tar and Coks	3148
7. Analytical Chemistry	3123	22. Petroleum, Lubricants, Asphalt	
8. Mineralogical and Geological ..	3440	and Wood Products	3144
9. Metallurgy and Metallography ..	3410	24. Cellulose and Paper	3151
10. Organic Chemistry	3437	24. Explosives and Explosions	3151
11. Biological Chemistry	3013	24. Oyes and Textile Chemistry ..	3159
12. Foods	3052	25. Paints, Varnishes and Resins ..	3153
13. General Industrial Chemistry ..	3057	27. Fats, Fatty Oils, Waxes and Soaps ..	3153
14. Water, Sewage and Sanitation ..	3101	28. Sugar, Starch and Gums	3151
14. Soils, Fertilizers and Agricultural		29. Leather and Glus	3164
Poisons	3109	30. Rubber and Allied Substances ..	3153

PATENTS

A statement giving information as to how to obtain patent specifications and drawings, both United States and foreign, is to be found on page i of *Chemical Abstracts*, 25, No. 11.

Published by THE AMERICAN CHEMICAL SOCIETY, which also publishes *The Journal of the American Chemical Society* and *Industrial and Engineering Chemistry*.

Communications for the editor should be addressed to E. J. Crane, Ohio State University, Columbus, Ohio.

Remittances for annual dues (\$15.00) in the American Chemical Society or for subscriptions to this *Journal* (\$7.50 per year), changes of address, orders for subscriptions or back numbers, claims for missing numbers and all matters intended for the *Proceedings* should be sent to the Secretary, Charles L. Parsons, Mills Bldg., Washington, D. C. Foreign postage is extra, being \$4.50 per year added to the dues of members residing outside the United States, its possessions and the countries of the Pan-American Union and \$1.50 per year for *Chemical Abstracts* alone, except that only one-half of these amounts is added for journals going to Canada. The countries of the Pan-American Union are Argentina, Bolivia, Brazil, Chile, Colombia, Costa Rica, Cuba, Dominican Republic, Ecuador, Guatemala, Mexico, Nicaragua, Paraguay, Peru, Republic of Honduras, Republic of Panama, Salvador, Spain and Uruguay.

The Council has voted that no claims will be allowed for copies of journals lost in the mails unless such claims are received within sixty days of the date of issue, and no claims will be allowed for issues lost as a result of insufficient notice of change of address. (Ten days advance notice required) "Missing from files" cannot be accepted as the reason for honoring a claim—Charles L. Parsons Secretary.

Published by the American Chemical Society, Easton, Pa. Secretary's Office: Mills Bldg., Washington, D. C. Entered as second-class matter at the Post Office, Easton, Pa. Acceptance for mailing at special rate of postage provided for in Section 1103, Act of November 3, 1917, authorized July 13, 1918. Mack Printing Company, Easton, Pa.

AUTHOR INDEX

[An asterisk (*) indicates that the reference is to an abstract of a patent]

- Abderhalden F, 2973,
2974
Abel E, 2844
Abelin I, 3039
Achenbach A., 2931
Ackerman J W, 2504
Ackland F W, 3143
Adam D B, 3118
*Adam W G, 3134
Adams, B A., 3104
3105
*Adams, E Wm
3119
Adcock F, 2880
Adell B, 2901
Adkins, H, 2978, 2985,
3133
*Adler L., G m b H.,
2883
Adler M., 3041
*Adler, O., 3013
*Adler, R., 3013
Adora A. N., 3048
Adonaeus L., 3034
Agnew, N V., 2938
2960
Ahlfeld F., 2940
Ahmad N., 2912
Ahmed, 3087
Ainsberg L. P., 2947
Aiyaz, C. V. R., 3029,
3030
*Akiyama T., 3136
*Aktiebolaget Bank-
firman Alfred Berg,
3148
*Aktiebolaget Sepa-
rator, 2881
*Akt.-Ges. Brown,
Boven & Cie, 2964
*Akt.-Ges. für Such-
stofflager, 3136
*Aktieselskabet Rasoul
Pictet & F Tharald-
sen, 3167
Albizziati, C. M., 3002
Alcázar J. N., 2960
*Albrecht H. W., 2963
Aleksandrovskii, V.,
3148
Allen, C. F. H., 2992
*Allen, H. I., 2927
Allen, L. A., 3093
*Allgemeine Elektrici-
tätsges., 2883,
2964
Alvarez C., 3051
Alyea H. N., 2904
*American Glue Co
3109
*Amoroso M., 2965
Anand C., 2906
Anderegg J., 3151
Anderson O., 3060
Ando N., 2909
*André H., 2926
Andrews A. J., 3143
Angeli B., 2924
*Angelucci O., 2882
Annetta, M., 2877
*Annex J., 3177
*Anode Rubber Co.,
Ltd., The, 3178
3199
Anselmino K. J., 3051
Antonini A., 3064
*Antonov V. N., 2882
Apelo F., 3063
Apperly F. L., 3046
Apushkin K. L., 3117
Archbutt S. L., 2953
*Archer R. S., 2965
Arend J., 2938
Arends B., 2921
d Argent, C. S., 3104
Anyama, T., 3034
Armandi C., 3023
Armstrong J. I., 3032
*Armstrong, P. A. E.,
2966
Arndt, W., 3099
*Arnold, E. F., 3164
*Arnold, H. R., 3012
Aron, M., 3078
Artero V., 3094
Arthur, M., 3063
Arzeni, R., 2933
Ashaba Y., 2990
Ashkanaty, M., 3099
*Ashkenazy, P., 3118
Aspin, F. W., 2912
*Attack, F. W., 3138
Aten A. H. W., 2887
Atkins, K., 3181
*Atwell H. V., 3152
Audreth L. F., 2925
Auer, L., 3156
Auermaier M., 2909
Auroch L., 3064
*Autogenwerk Suis-
s, G. m. b. H., 2883,
Auzina O., 3092
Avery, R. C., 3054
Aviles L., 3025
*Axelrod A., 3097
*Ayerst D. B., 2927
Ayyaz, P. R., 2971
3007
B. H., 3174
Baehack E., 3030
Bacq Z. M., 3082
Bader D., 2884
Badische Maschinen-
fabrik & Eisengies-
serei vormals G
Schold, 2964
Badstuber W., 2880
*Baer J., 3199
Bagdas G., 3027
Bakla A., 3163
Bahr H., 2932
Banea H., 2929
*Bakelite Corp., 3135
3145, 3164, 3185
3186
*Bakelite G. m. b. H.
3185
*Baker & Co. Inc.
2927
*Baker H. H., 3137
Balabucha - Popcova
V., 3124
*Baland H., 3131
Balas F., 2990
Baldwin F. O., 3104
*Balke C. W., 2927
Baltman F., 3018
Banerjee A. K., 2950
Bancroft D., 2911
*Bannister, W. J.,
3014
Bansall J. W., 3098
Banda, A. G., 3179
Barat C., 2908, 2976
Barattini G., 3128
Barbour, H. G., 3081
Barbour, J. P., 3166
Bartholomew, J. C., 2988
Barth, G. H., 3020
Barrot, H., 2884
Barker, L. B., 3106
*Barker, W. T., Jr.,
3144
*Barnes C. J., 3169
Barnes D. J., 3036
Barnett, E. de B.,
2994
Barone, V. G., 3037,
3085
Barrenbrecht H. K.,
3087
Barrett E. G. V., 2914
Barry, C., 2932
Bart, 2893
Barthel M., 2909
*Bartlett, Hayward
Co. The, 2882
*Barthug, F., 2884
3153
*Baschke, E., 3194
*Bateheller, C., 3147,
3148
Batson R. O. C., 2957
Battagay M., 2942
Batutkat T., 3146
Bauer H., 2910, 3081
Bauer K. H., 3124
Bauer W., 3079
Baum R., 2924
*Baumann M. C.,
3179
Baumgarten, K., 2962
Baxter, G. P., 2885
Baylis J. R., 3103
Bayton L. E., 3046
Bayo C. P., 3090
Beccari E., 3049
Beck J. W., 2910
Beck P., 2952
*Beck W., 2964
*Becker J., 3154
Becker J. A., 2925
Becker Rose H., 2884
Bedel C., 2903
*Beekley J. S., 3133
*Begover M., 3160
Behr J., 3070
*Beil Adel J. A., 2966
Belikova M. I., 3110
Bell J., 3093
Belkin, V. P., 3115
*Belunski, S. V., 2882
Benade, J. M., 2912
Bencsik F., 3087
Bender, E., 3106
Benford, F., 2919
Bennett D., 3073
*Bennett, L. J., 3097
Bennett, R. D., 2887
Bennett, K., 2903
*Benson F., 3175
*Bergdorfer Eisen-
werk A. G. Antra-
werke, 3096
*Bergelt C., 3190
Berger, I., 3020
von Bergkamp, E. S.,
2979
Bergmann E., 2920
*Bergmann, E., 2926
Bergmann, M., 2977
Bergstrom P. W., 3001
Berl P., 2979, 2983
*Berliner Hartpulver
Ges., m. b. H., 2966
Bernardi, A., 2937
*Berry, E. E., 3170
Berthelot, A., 3025
Beschorner, M. A.,
3139
Besonova A., 3026
*Beste K., 2884
*Betout-Ges. m. b. H.,
3147
Bevenders, R. G., 2985
Beyer, H., 3008
Bhalia R. L., 2961
Bhatia, S. L., 2887
Bhatnagar, S. S., 2887,
2920
Bhattacharya, R.,
2971
Biazio, R., 2939, 3142
Bichowskaya, A., 3080
Bichowsky, F. R., 2997
Bielmann E., 2971
Bielow P. A., 2992
Bier A., 3051
Bier O. G., 3055
Bierrett, G., 3166
Berry, H., 3016
Bigwood, E. J., 2900
Bihlmaier, K., 2925
Bikerman J. J., 2895
Birk A., 3063
*Bilner K. P., 3147
Bills M., 2929
Bills W., 2885
Binaeh, R., 3050
Bina A., 3021
Bischhoff P., 3081, 3089
Bishop, L. R., 3120
3121
Biswas S. L., 2949
Bjeler M. E., 2907
*Black J. A., 2881
Black J. H., 3056
Blackie J. J., 3129
Blair C. A., 3195
Blair, C. W. S., 2890
Blake G. S., 2946
Blanchard, L., 2975
Blas, L., 3103
Blaschko, H., 3046
Blau M., 2914
*Blecker, W. F., 3187
Bliebaum, L., 2923

- *Bley, G. 3178
 Bleyberg, W. 3186
 *Bleyer, C. 3154
 *Blomqvist, G. M. J. 3109
 Bloom, N. 3081
 Blum W. 2960
 *Blythen, S. 3183
 de Bobest B. 3132
 Bobrov, P. 3183
 Bobrov, P. A. 3156
 Bodansky A. 3091
 Bodenlander H. G. 3181
 Bodenstein M. 2904
 Bodifors S. 2921
 Boedeker F. 2981
 *Böhme, H. T. A.-G. 3133
 *Boehring, C. P. & Soehne G. m. b. H. 3137
 Böschütz, G. 3106
 Bohn, H. 3071
 *Bohrer, R. E. 3170
 Bolam, T. R. 2909
 Bolander, L. M. 2983
 Bolshovnikov, N. F. 2958
 Bollinger, G. M. 2902
 *Bomke, H. 3136
 Bonanni, A. 3084
 Bonaretti, A. W. 2958
 Bonate, P. 2947
 Bond W. R. 3084
 Bondelli, M. 3025
 Bone, W. A. 3068, 3171, 3173
 Bonniel, R. 3116
 Booth G. 3038
 Borchardt, W. 3067
 *Borchardt, W. O. 2963
 Borchardt, W. O. 3100
 *Borg, O. 3123
 Born, M. 2907
 *Bornkessel & Co. m. b. H. 3111
 Bornskaya D. P. 3057, 3059
 Bose, D. M. 2922
 Bose, N. N. 2917
 Bose, P. K. 3002
 Bossanyi, I. 2903
 Bottger, T. 3018
 Bouchaert, J. J. 3081
 *Boughton, W. A. 3139
 Bousset, E. 3079
 Bourguet M. 2980
 Bourne, M. C. 3056
 Bourquin, H. 3061
 Boyahn, E. 3104
 Boyd, T. A. 3156
 *Boyle, W. A. 3100
 Boynton, P. 3103
 Bracalosi, L. 3039
 Bradfield, R. 2896
 Bradley, J. 2957
 Bradley, R. S. 2891
 Bradshaw, J. H. D. 2937
 Brady, M. J. 3036
 Brandendörfer W. 3074
 Branczan J. A. 3099
 Brass W. 3113
 Brauch's P. 3019
 Braun H. J. 3140
 Brown K. 3087
 Brauns, D. H. 2977
 Brannan F. 2978
 Bravo G. A. 3173 3191
 *Breda J. 2892
 Bredlin H. 3126
 *Brecht G. 3131
 Brems A. 3087
 *Brenk H. 3118
 Brenner D. 3073
 Brentano C. 3031 3045
 Brégnière L. 3114
 Brewer A. K. 2923
 Brezovsky F. 3086
 Bridel M. 3097
 Briles T. C. 2909
 Brilman P. W. 2882 2906
 Brileu T. 3079
 Briggs, D. B. 2909
 Briggis H. 3119
 Brighton T. B. 3132
 Briley, G. 3184
 Bril's IL. 3154
 Brinley, G. W. 2916
 Brink C. 3116
 Briscoe H. V. A. 2933
 Britton H. T. S. 2903
 Britton, S. W. 3013
 Brocher, J. F. W. 3071
 Brodtkorb, P. 3132
 Brodovich, K. F. 2913
 Broniewski, W. 2958
 Bronk, D. W. 3048
 *Bronn J. I. 3153
 *Brooks, C. 3097
 Brooks M. M. 2983
 Brooks, S. C. 3033
 Brot, M. 3183
 Brown, A. 3031
 *Brown, J. M. 3195
 *Brown J. P. 3170
 Brown, M. 3018 3083
 Brown, W. L. 2913
 Browne, M. W. 3092
 Brownell, K. A. 3044
 Bruce, W. F. 3000
 Bruch E. 3040
 v. Bruchhausen F. 2983
 *von Bruck, H. 2964
 Brünner, K. 2937
 Bruhat, G. 2920
 *Bruhn, B. 3116
 Bruhn's G. 2939
 de Bruin, G. 3163
 Bruns, H. R. 2899
 Brull, L. 3036
 *Brunel, C. 2883
 *Brunelle, G. A. 2983
 Brunelli, L. 3077
 Brunetti, R. 2910
 Brunius, F. 3052
 Brunton, C. E. 3046
 Bruylants, P. 2969
 Buchanan G. H. 2932
 Buchholz C. 3070
 Buchwald, K. W. 3081
 Buck J. S. 2992
 Budakov P. P. 3112
 Bugbee L. P. 3012
 Bahariwala 2903
 Busson H. 2922
 Bust D. M. 3098
 BuPonok H. 2998
 Buxley R. 2930
 Bull L. 2899
 Blumming G. 3128
 Buzzer W. 2929
 Burger M. 3031 3071
 Burgess W. M. 2901
 Buschheim F. 3011
 *Burke F. 3100
 Burkhardt A. 2960
 Burke J. D. 2951
 *Burman A. S. 2963
 Burrell F. 2883
 Burstin H. 2879
 Burton P. P. 2887
 Burtcher J. 3011
 Busch G. 3181
 Buschinger H. 3179
 Busenbarg F. H. 3196
 Busquet H. 3079
 Buswell A. M. 3107
 Busz K. 2913
 *Butchart W. A. 2943
 Butler, B. S. 2918
 Butler P. W. W. 3105
 Butler, T. 3101
 Butz H. 3018
 Buxtorf P. 2990
 Buzanovskii, I. S. 3053
 Caldwell, F. H. 3161
 *Caldwell S. M. 3199
 Caille, A. 3161
 Caley, G. P. 3189
 Camacho M. A. 3079
 *Cambron A. 3011
 Canulla, S. 3093
 Campbell J. M. 3156
 Canals, P. 2906
 *Cannon, H. H. 3159
 Cantegrel F. 3036
 *Caplan S. 3154
 Cario R. 3088
 Carlos A. S. 3118
 Carlson A. J. 3011
 *Carnegie, D. Jr. 3184
 *Caro N. 3099, 3134
 Carpenter, D. C. 3019
 Carpenter, P. L. 3188
 Carroll J. A. 2880
 Carruthers A. 3029
 *Carter, F. C. 2884
 Cartwright, C. H. 2880
 *Carugh, A. 3131
 Castagne, E. 3034
 Castille, A. 2922
 Caverly, W. R. 2939
 Cavatone, A. 2914
 de Cohn, M. G. 2937
 Cerchez, M. V. 2971
 Chabrol F. 3078
 Chaklader M. N. 2976 2999
 Chakraborty, S. N. 2977
 Chakravarti G. C. 2998 2999
 Chalange D. 2981
 Chamberlain P. N. 3083
 Chang H. 3086
 Channasappa, H. M. 2898
 Chapman A. W. 3106
 Chamaus K. 3181
 Chatterjee S. K. 2942
 Chatterji A. C. 2999 2992
 Chaudhuri, H. 3028
 *Chevance L. 3153
 de Chaves P. 2935
 *Chem Fab Grünau Landkloß & Meyer A. G. 3196
 *Chem Fab von Heyd A. G. 3016
 *Chemische Fabrik Kurt Albert G. m. b. H. 3185
 *Chemische Fabrik vorm. Sandoz, 3128, 3173
 Chen, T. T. 3038
 Chernikov, 3092
 Chernikov, R. S. 3037
 *Cherry, O. A. 3136
 Chesson R. 3183
 *Chesny H. H. 2882
 Chevenard, P. 2958
 Chhiber H. L. 2956
 Chikashiro, M. 2907
 Chinn P. 3103
 Chmelof F. 3192
 Chmura, M. 3125
 Chokanna N. G. 2925
 Choknoky, L. 3029
 *Christensen, H. K. 3097
 Chung H. L. 3095
 Chvorinov, N. 2925
 Cianci V. 3128
 *Ciando, H. 3131
 *"Carle Werke" Josef Lorenz & Co. 3184
 Ciprian C. 3049
 Citron H. 3081
 Quin W. 3091
 Clafin, A. A. 3195
 Clark, A. E. 3102
 Clark, E. P. 2987
 *Clark, F. M. 3160
 Clark, G. A. 3083
 Claude, A. 3056
 *Claus, G. 3134
 Clavera J. M. 3121
 *Clemens R. 2883
 Clemm, G. R. 3006, 3007
 *Clerc, R. 3013
 Cleveland, H. B. 3101
 *Clifford, W. B. 2883
 *Coberty, C. J. 2964

- *Collier M 2884
 *Code G A 2963.
 Colman V 2996
 Coghill W H 2940
 Cohen F L 2942
 Cohen W E, 3164.
 Cohn M M 3106
 Cohn W M 3141
 *Cole S S 3145
 *Coleman C 3199
 *Coleman R I 3100.
 Cole J 3043.
 *Colorado Iron Works
 Co 2963.
 Colwell, A R 3080
 Colwell D L 2958.
 Comel M 3034.
 *Compagnie Francaise
 pour l'exploitation
 des procédés Thom-
 son Houston 2968.
 Compagnie générale d'
 électricité 2926.
 *Comptoir des textiles
 artificiels (Soc.
 anon.) 3164.
 Constant J R., 2991.
 Cone R. M., 2945.
 *Conner, A., 2964
 *Concordia - Bergbau
 A-G, 3153.
 de Courcy, E. H.,
 3098.
 Connolly, J. L., 3105.
 Connor, R. 3133.
 Conrad C M 3095
 *Consortium für elek-
 trochemische Indus-
 trie G m. b H.,
 3018.
 *Continental Parker
 2966
 Cooke F C., 3148.
 Cooke, S. B., 2960.
 *Cooper C 2982
 *Cooper H. S. 2943.
 *Cope F T 2964.
 Cordier R., 3017
 Cotti, C. F 3051
 Cori, G T, 3081
 Corkhill A B., 3046
 3043.
 *Cornell F G., Jr.,
 3097
 Correy G W 3042
 *Corning Glass Works
 3144
 *Cortwell V O 3144
 Correa, L M 3046.
 Cossin F 3087
 Costa A 3037 3052
 Cotton F H 3142
 de Coulon A 3064
 *Coutant J G 2943
 Cowan, H. MCN
 2915.
 Coward K. H 3035.
 Cowles P B 3026
 Cox C. R. 3104
 Crafts A. S. 3033
 Craig L C 2997
 *Craig P H 2983.
 *Craig R. F 3177
 Craik J 3163.
 Crandall L. A. Jr.,
 3089
 Cranes D F, 3133.
 Crepez, E 2906 2955.
 Crespi M 2904
 Crislin B 3125.
 Crookman, L. G.,
 3106
 *Cross, W M 2554.
 Cossard J 3093
 Cummins H A 3004
 *Cuningham, O D.,
 2983.
 Curio G 2944
 Curran H R., 3026
 Curreber Y 2941
 Cury V., 2981 2903.
 *Cusack R 3165.
 *Cuspi, T., 2943.
 Delach, H., 3151
 D'Achiardi G 2941
 Daddwell H F, 3164.
 *Daubs Käpyd Sei
 yaku Kabushiki
 Kaisha 3116
 Dale H H 3046.
 Dalling T 3064
 Daly, S F 2946
 D'Amour F E 3044
 Damm E, 3071
 Dangeard P., 3030
 Danelopolis D., 3072.
 Danelop F 2904
 *Darling F R 3000
 Darton, 3194
 Das-Gupta P N.,
 2934.
 Datta, P C. 2945
 Daubau Dehls J.,
 2946.
 *Dauid H B., 3016
 Davies J G., 3033
 Davies, R. J 2940.
 Davies W C., 2970.
 Davis, D S., 2989
 Davis H L 2904
 Davis I A., 3197
 Dawson A. B., 3064.
 *Day R B., 3199
 *Deasendy R M
 3014
 Deb S C. 2920
 Debas, G 3047
 *DeCew J A 3170
 *Deconclois E 3175
 DeJek, J 3193
 De Fau R 3121
 Defontex, P 294.
 Degorge, H 3000
 De Graaf 3126
 De Gonds C 2949
 De Hooy L 2555
 Dekker L 3111
 *Delaney M E 3127
 3128.
 Delcourt Bernard E
 3013
 Delgado R 3087
 *Delpech J 3165
 Delucy 3114
 De Michelis, U., 3062.
 Demina, E 2924.
 Demson G H. 2957
 Demville, L., 2942
 *Depasse, E., 2982
 3099
 Depew, H A., 3197
 *Derby I H 2903.
 Derra J., 3087
 Derrbach, W., 3125.
 Deval, B N., 2900
 Deschamps R., 3021
 Desmaroux, 2927 3163.
 Desvergne, L., 2940.
 Detorre J., 3116
 *Deutsche Bergbau-G
 für Kohle-und Er-
 dölchemie, 2942.
 *Deutsche Gas- und
 Licht Auer-G m. b.
 H., 2943 3134
 *Deutsche Gold und
 Silber Schmelzwerk
 vorm Reichert 3160
 Devine J M., 2961
 *De Vries, E P 2968.
 Dewey W A 3072
 De Wild A M 3143.
 Derram S 3124
 Dhar N. R., 2921,
 2977 3029
 Dhérif C., 3017
 *Diamond J J., 3176.
 Dice C M., 3132.
 Dieckmann G., 3132.
 Diego V., 3157
 Diehl K., 2997
 *Diescher S. E., 2964
 Dietrich E O., 3193.
 Dietzle R. A., 3154
 Dietzle, H., 2968
 *Diets, E., 2976.
 *Diets, V., 3139
 Di Giama, J., 2903.
 van Dyk J A 3157
 Dickens, P., 2979.
 Dill D B 3043.
 Ding, M. 3075.
 Dirken M. N. J.,
 3047
 Di Stefano F 3125.
 Dittmer M 3174.
 Dworkin F F 2978.
 Dixon A A 2920
 Dixon M 3019
 Dlouh B 3193.
 *Dobón J 3160
 Docherty J G 2929
 Dorell E G 3113
 D'Oring G 3172
 Dorr E 3039
 Dohognt, A 3192.
 Dohrow J A 3121
 *Dolch M 3145
 Dole M 2902
 Dolko Dobrovodskij
 V V 2044
 Dolphin J R 3094
 Donan J 2979
 Donovan H 3073.
 Dopolov P K. 3124
 Dornte R W 2557
 Doss K G 2908.
 Dostál L 3193
 Downes J R. 3107
 Drake, T G H. 3034.
 Drake W E 3127
 Drfan, G., 3180
 Drechsel W., 3164
 Dregman, M., 3087.
 Drees G. M., 2919
 *Dreier, C., 3167.
 3165, 3169, 3176,
 3177.
 *Dreyfus, H., 3011
 3167, 3169, 3177
 Dreyfus, K., 3052.
 Dreyssing C., 3117
 Drier R. W., 2906.
 Drisker C. K., 3034.
 Drogis L., 3196.
 *Drouilly, E., 2927
 Drury, A. N., 3063.
 Drutman, D V., 3114.
 *Dubba, C. P., 3155
 3159
 Dubey, V S., 2945
 Dubois, E., 2944
 Dubois, M., 2929
 Dubois, R., 3039
 *Duckham A McD.,
 2943 3153.
 Duff H., 2971
 Duffey J., 2910.
 *Duffield F L., 3153.
 *Dufraisse, A., 2962
 Dufraisse, C., 3171
 Duke-Elder, F M.,
 3043.
 Duke-Elder, W. S.,
 3043.
 Dukes, H H., 3048
 *Dumas, M., 2982.
 *Dunbar T L., 3169
 *Dunham, G W.,
 2981
 Dunlop J F 2959.
 *Dunlop Rubber Co.,
 Ltd 3178, 3199
 Dunning H. B.
 2901
 Dunning P 3127
 Dunstan A E 3103.
 Dupare L., 2934
 Dupont, G 3149.
 *Du Pont de Nemours,
 E. I. & Co., 3133,
 3172
 Dutt N C. 2977.
 Dutt S K. 2918.
 Durnette W J A.,
 3151
 Dwarik L 3113
 Dworkin S. 3092
 Dyer F J 3035.
 *Dyer H H., 3139
 Eadie H 3065.
 *Earl A R., 3159
 Eastman E D., 2908.
 Eaton M D., Jr.,
 3076
 Eckart O 3133.
 *Edelmann L., 2981
 Eder, R., 3123.
 Ederle, W., 3072
 *Edmonds, W J.,
 3153.
 Edwards H. D., 3095
 Edwards H T., 3043.
 Egerton A., 3146

- [illegible]

- Gomer F S 3191
 Good W 2487
 Goodman A II, 3106
 Goodyear F II, 3173
 Goots R 2977
 Gorsalevich D K, 2931
 Gordon W S, 3064
 Gore B 2931
 Gornis P 3064
 Goss W G 3094
 Gosper B 2941, 2942
 *Goudakristvennaya
 Suureno Kram
 naya Fabrika Goss
 darstvennaya Terats
 Leningrada
 Shvenoi Promysh
 lennosti Leningra
 doderbda 3177
 Goto L 3002
 Goto T, 3114
 Gowen C II, 3028
 Goy, S, 3112
 Gósony L (Ludwig)
 3054
 Graber, L. F., 3033
 Grabow C, 3054
 Grabus M, 3109
 Gradmann II, 3109
 Gräf, R., 2999 3000
 Graf W, 3074
 Grade V, 3130
 Grabins A, 3017
 Graham C J, 3099
 Gramenick M I, 3076
 *Grandenhe R II, 3097
 Grasdare, A., 3055
 Grasse L I, 2949
 Grasse, O A, 2914
 Graver, L F, 3033
 *Graves R R, 3098
 Gray, J W, 3026
 Gray L II, 2913
 Grebel A, 3150
 Greco A T, 3142
 Green M F, 3061
 *Greenwalt, J E, 2963
 Greenbaum S S, 3066
 Greene C II, 2883
 Greenhill A W, 3117
 Gregory, J C, 2909
 Greg M E, 2973
 *Greengr R 3147
 *Greene, H., 3170
 Grishanov E, 3052
 Gries 3071
 Griesel F 2924
 Griffin R C 3166
 Grifth R C 2962
 Grigast A 3021
 Grigorev P N 3027
 *Grinder E 3178
 Grimes, H. M. 3042
 Grimes M 3094
 *Groff F 3128
 Grolman A 3043
 *Gronau E 2930
 *Gruhenholcim
 pragerung G m
 b II 3143
 Gruchl, H L 3065
 Gruen, R, 3141
 Grünwald M, 3109
 Grüss J 3030
 *Grud P, 3096
 *Guadagnani A, 3194
 Guarnieri P, 3093
 *Gubelmann I, 3178
 *Guernsey, F II, 2940
 *Gugrenbühl, M A, 2452
 Guha P C, 2976, 2977 2990
 Guillet L, 2944
 *Gumma und Balata
 Werke "Metador"
 A-G, 3109
 Gunde M 3025
 Gupta, B M 2976
 Curreich J I 3049
 Gurevich L M 3117
 Gurevits J 3053
 Gusev S F, 3113
 Gustavson R G, 3044 3099
 Guthrie J M, 3172
 Gutwiler R, 3191
 Grylls Z, 2962
 *Haar W & Co, 2883
 Haas A, 2923
 Haas W, 3123, 3174
 Hadrowski A, 3070
 Häusler, E F, 3049
 Haeremans A 2982
 *Häcker K 3176
 Hahn S II, 3198
 Hais J T, 3111
 Hattiger M 2880
 Hattner W, 3035
 *Hefner P, 3012
 Hale White Wm, 3130
 Hall A J, 3173
 *Hall E J, 2884
 Hall, N F, 2901
 Haller, J, 2879
 Hallett, R L, 3160
 *Halstead R T, 3167
 Halvorsen H O 3108
 Hamada, H 2917
 *Hamboyer S 3160
 Hamano H 3065
 *Hamed C A C 3165
 *Hamel F 3168
 Hamid M A 2904
 Hamilton R M P 2951
 *Hammermill Paper
 Co 3194
 Hammond C F 3131
 Hamova J 3192
 Hamsh A 3016
 *Hansen C I 3167
 Hanns A 3079
 *Hannestische Möbel
 anwerke C G 3195
 Hansen F 3113
 Hansen K F W 2989
 Hansen M 2958
 Hanson, D, 2953 2961
 *Hanson F R, 3138
 Hanson, F, 3112
 Happe 3071
 Hara, H., 3042
 Hara K, 3057
 *Harding, A T, 3160
 Harding V J, 3050
 Hardy, R L, 2961
 *Hardy, T W, Jr., 2945
 Hartman W D, 2911
 Harley, D, 3129
 *Harris, O, 2964
 Harrington E L, 2913 2914
 Harris F C 2893
 Harris, L J, 3050
 Harrison A W C, 3181
 *Harron J A, 3178
 v Harnel II, 2885
 *Harter, I, 3144
 *Hartford, F M, 3144
 Hartman, F A, 3044
 Hartman, W A, 3044
 Hartmann E, 3054
 *Hartmann, R A, 3130
 Hartmann W, 3099
 Hartung W II, 3090
 Haas K II, 2951
 Haas, R., 2954
 Hasegawa, T, 3067
 Haslam C S 3180
 Hata C, 2984
 Hatano T, 3034
 Hatch F II, 2930
 Hatcher, M F, 3102
 Hatcher R A, 3130
 Hatos, G 2936
 Haus Haher F L, 3197
 von Hausen S 3030
 Hauser E A 3197
 *Hause P 3160
 Haven F L, 3066
 Havens L C 3056
 Hawley J T 3106
 Hawley L F 3163
 Hayakawa, K 2973
 Hayama N 2919
 Hayashi T 2995
 *Hayden R 3147
 Heub C M 2956
 Heber K 3124
 Hebbelberg II, 3181
 *Heberlein & Co A
 G 3177
 Hecht K 2917
 *Hedberg C W J, 2923
 Heffler A 3092
 *Held J B, 3158
 *Heddelberger 31 3027
 Heiduschka, A, 3093
 Heigl A, 2950
 Heiman F, 3053
 Heimer A 2918
 Heinemann, A C, 3181
 *Heinrich C, 3168
 Hecker N, 2933
 Heffernich, B, 2973 2977
 Heimer, O M, 3056
 Heimbolt, H F, 3091
 *Hemingway, H J, 3155
 Henderson, C W 2916 2950
 *Henglein P A, 3134
 Henne, A L, 2967
 Henze, M, 3040
 *Heraeus - Vacuum
 schmelze, A-G, 2928
 Herbert, F K, 3056
 Herbracht, H, 3120
 Hennig G C, 3193
 Herkel, W, 3031
 Hermann, S, 3068 3071
 *Herrmann, S. 3129
 Herold, W., 2921
 Herndon L K, 3108
 Herold R D, 3023
 Hertel, E, 2927, 2933
 Hertzel, E A, 3143
 Hersberg, E, 2989
 Hersfeld, E., 3027 3073
 Herzog R O, 2900
 Herzog W, 2940
 Hess, W F, 3167
 Heutenich, H, 3140
 Heulser W, 3092
 *Heuser, R V, 3200
 Hubbard, W A, 2962
 Hubbert, H, 2973 2978
 *Hubl K, 2928
 Hücking O A, 2944
 Hicks J F G, 2961
 Hiedemann E, 2993
 Hietcher A 3141
 Hildebrandt, F, 3071
 *Hilger, G, 3100
 *Hilgers, E., 3144
 *Hill, B I, 2853
 Hill C A, 3165
 Hill M, 3048
 *Hill S B N, 2982
 Hill, S E, 3039
 Hind, H L, 3121
 Hinteregger, P, 3070
 Hirata, H, 2891
 Hirose H, 3032
 Hirsch, H, 3141
 *Hirsch Kaper- und
 Messungswerke, A-G, 2927, 2928
 Hirsch, W C, 2955
 Hirsfeld L, 3058
 Hiseck, D J, 3114
 *Hittner, H P, 3144
 Hitzig, W, 3089
 Hixon, R. M., 2997
 Hochberg, R B, 3052
 Hoder, F, 3026
 Hodge-Smith T, 2943
 Hodge W W, 3105
 Hodson G A 3141
 *Höfken M, G m. b
 H., 3169

- *Hinkende, G. H. B.
H., 3167
Hilte, R. 2932
Hilte, F. 2930
Hilte, J. 2947
Hilte, H. 2932
Hoffenreich F. (Franz)
3034
Hoffmann, F., 3031
Hoffmann G. 3151
Hoffmann C. H. 2979
Hobbs J. 3034
Hobbs D. 3154
Hodder, L. 3154
Hole, A. S. 3154
Hoard A. 2927
Hofer K. 2943
Holman R. 3031
Holmes, H. N. 2924
Holmvard F. J. 2931
Holschneider F. B.
3033
Holten C., 2932, 2937
Holton, E. C., 3154
Holweg, E. 2934
*Holschneider A.-G.,
3194
*Holworth, E. H.,
2964
Holzer J., 2911
*Hower, W. C., 2963
Hooper, G. K., 3104
Hoover, J. R., 3197
*Hopkins, H. H., 3144
Horeka, R., 3171
*Hori, S., 3139
Horikiri, M., 2931
Horn, C., 3034
Horn, G., 2916
*Hornbuckle, W. P.,
3174
*Horn, J. M., 3117
*Horsfield, R. T., 3131
Hoesters, H., 2923,
3031
Hoshikuni M., 3034
Hoschkins A. G., 2924
Hou, H. C., 2937
Houw A. G., 3133
*Howard R. H., 2964
Howe, R. F., 2919
*Howe, S. P., 3137,
Houw, W., 2933
*Hoyer, H., 3194
*Hoyer, P., 3194
Hort, S. L., 2921
*Humpfer, F. W.,
3137,
Humpfer, M., 3197,
*Huff, L. C., 3154
Hübner, E., 2914
*Humboldt - Deuts
motoren, A.-G.,
3161
Hume-Rothery, W.,
2929
Humphrey, H. A.,
2934
*Hunady, L., 3134
*Hupert, A., 2937
Hurd C. D., 2922
Hurst, A. F., 2961,
Hurst, J. R., 2933,
*Husemann, E., 3133,
Hude R. R. 3034
Hudel A. 2932
Hie W. S. 2992
*H. G. Fortbemaßung
A. G. 2942, 2927,
2943, 2911, 2912,
2913, 2914, 2915,
2916, 2919, 3100,
3119, 3123, 3125,
3133, 3134, 3135,
3137, 3138, 3139,
3142, 3143, 3149,
3160, 3167, 3164,
3169, 3174, 3175,
3176, 3177, 3178,
3184, 3185, 3186,
3193, 3199, 3200,
*Hwa K. 3137
*Hwa T. 3133
Huskins T. 2923
*Hwa G. 2943
Imbue C. C. 2945,
2933
Imhof, B. G., 2933,
2943, 2941
*Imperial Chemical In
dustries Ltd., 2914
3173
Imre L., 2927
Isaba, R. 2927
Isagaki, G., 3109
Isagaki, R. A. 3033
Ingensol, L. R., 2931
Ischem, H. 3103
*Institute of Physical
and Chemical Re
search, 3133, 3174
*Institut für physikal
ische Grenzschicht
Mechanik, 2927
Isami, S., Jr., 3131
Isawa, E. E., 3123
Irvine L., 2944
Irvine, J. C., 2945
Ischikawa M. P.,
2933
Ischikawa, M., 3117
Ischikawa, M., 2941
Ischikawa, T., 2944
*Isikawa, S. 3137
Isip, M. C. G.
2944
Issei, P. S. 2933
Isosaka, G., 3141
Ito, R., 3127
Ivanenko, D. 3103
Ivanov D., 2945, 2947
Ivanov, S., 3146
*Ivanova, V., 3103
Iwaki, K., 2934
Iwata, H., 2934
Iwengar M. S., 2933
Iyer M. P. V., 2934,
2933, 2934
Izari, J., 3171
Jacobs, F., 3194
Jod A. J., 2933
*Jaeck, A. O., 3133,
3133
Jaeschke, A., 3141,
3133
Jaffe, H. L., 2991
Jaffe K., 2931
Jain, M. S. 2901
James, F. M. 2934
Jankov 2931
Janet C. 2912
Jappet, A. 2931
Jarno, L. 2940
*Janvier G. F. 3134
Jans, J. H. 2911
Janssen, L., 2912
*Janssen (Janssen)
Schott & Co. 3111
Jankins, S. H. 3173
Jankins H. 3173
*Jansen J. J. 2943
Janssen, A. 2929
Janssen C. 2927
*Janssen, I. 2924
Jans, W. R. 2932
*Janssen A. 2934
Jans, W. 3102
Janssen F. B. 2914
Jans, A. 2913
Jans, D. S. 2914
Jankowski, A. 2934
Jankowski, F. M. 3103
Jankowski, F. F. 2934
Jankowski, T. B. 2931
Jankowski, W. C. 2931
Jankowski, H. P., 3131
Jans, H. S. 2914, 2913
Jones, A. I. 3129
Jones C. L. 3123
Jones, E. 2934
Jones, G. 2932
*Jones, H. L., 3174
*Jones, R. B., 3179
Jones, T. G. H. 3124
Jones, W. J. 2931
de Jong H. G. R.
2933
Jons, V. 3133
Jordan, H. 2913
Joshi, K. A., 2941,
2943
Joshi, V. G. 2927
*Jost, F., 3133
Jost, R. 3116
*Jost, R. E. 2922
Jung, H. 2943
Jungkunt, R. 3124
Juhn, M., 2944
*Jurech, H. 3104
Jurech P. J. 2931
Justin Moeller 3162
Kaas, A. I. W., 2931
*Kardier H., 3160
Karpf, A., 2933
Karpf, G., 2934
Kafuku, K., 2931,
3123
Kallias, J. G., 3103
Kallias, M., 3133
Kamash, R., 2933
Kamash, R., 2933
Kamash, K., 3141
Kamash, 3192
Kammer, F., 2933
Kamsh, J. D., 3164
Kane R. E., 2934
Kane, H. 3173
*Kansai Paint Kabu
shi Kaisha 2913
Kappelmeyer C. F. A.,
3142
Kaplanow R., 2943
*Karl, I. 2944
Karl, R. 2913
Karna, N. S. 3144
Karnish, L. 2931
Karnish, G. S. 3103
Kashira, S. 3171
*Kastner R. 3171
*Katsura Denko
Kero Kabushiki
Kaisha, 3119
Kata, G. 2974
*Kato, H. 3174
Kato, R. 3173
Kato, A. 2919, 2923
Katsura, C. G., 3157
*Kauf, O. 2944
*Kauf, F. 3112
Kawa, K. 3124
Kawar, W. 2933
Kawar, E. P. W.
3194
Kee, B. H. 2937
Kee, G. L. 2942
Kee, W. H. 2944
Kehoe, R. A. 2941
*Kell, O. 2933
Kell, A. M. 3132
*Kell, F. C. Jr.,
2941
*Kell, K. 3133
*Kell, G., 2927
Kell, E. 3132
*Kell, F. C. 2924
*Kell, M. W., 2944
*Kell, E., 2924
*Kemper, H., 2939
Kemp, J. D., 2947
Kemp, L. W., 2943
Kemp, R. 3129
Kendall, F. E. 2927
Kemp, V. C. E.,
2941
Kemp, L., 2934
*Kemp, G., 3114
*Kerr, J. G., 3174
Kerr, R., 2942
Kerr, H. E., 2937
Kerr, J. R. C.,
3104
Ketelaar, J. A. A.,
2943
Kew, C., 3144
Ker, K. M., 2933
*Kharasch, M. S.,
3123
Kharasch, V. N.
3113
Kharasch, M.
O., 2933
Kharasch, S. R., 2911
2913
Kharasch, A. L., 2933
Kharasch, P. K., 2913
Kharasch, T., 3119
*Kikui, T. 3133
Kikuchi, R. A., 2944
Kikuchi, K., 3127
Kimura, K., 2941
Kin, K., 2930
Kindler, K., 2941
Kindler, E., 3141

- King A. M. 2891
 King A. T. 3174
 Kirmann, A. 3016
 de Kiss A. 2903
 Kietler S. S. 2896
 *Kita G. 3178
 Kiuti M. 2916
 Klatt F. 3126
 *Kleimncker P. 3006
 Klemm, L. 2886
 Klemm W. 2886
 2891
 Klemp W. 3132
 Khager P. 2929
 Kluke J. 3017
 Kluijver A. J. 2884
 Klyuyeva N. G. 3026
 Knapp O. 3140
 Kneuer A. 3006
 Knight W. A. 3130
 Knott, P. A. 3061
 Knowles C. R. 3103
 *Knowlton T. E. 3148
 Kobayashi, Y. 3091
 Kobe K. A. 2932, 3162
 Kober, K. 3072
 Koch, 3192
 Koch, H. J. 2939
 Koch, S. 2940 2942
 Kochmann M. 3072
 Koenig, J. E. 2903
 *Kodak Patk. 3167
 Kodama T. 2939 2971
 Köhler K. 2951
 Koepfen S. 3088
 Koers J. H. 2906
 Koder, L. 3125
 *Kobler, A. M. 3144
 Kohler, E. P. 3000
 Kobb, S. J. 3098
 Kohlachia W. 3171
 *Kohorn O. & Co 3168
 *Kohul, S. 3146
 *Kokmao M. I. 3147
 *Kokolenbau und Geyerwertung A. G. 3154
 Knibath W. 3034
 *Koller K. 3154
 *Koller M. M. 3096
 Kolodenezky, J. 3058
 *Kolozsy F. 3138
 Kolthoff I. M. 2901, 2909
 Komatrubara H. 2891
 Komaya, K. 3058
 Komay E. 3093
 Kondo T. 3090
 Konek F. (Fritz) 3004 3005
 Kopecky O. 3192
 *Koppere A. 3154
 *Koppers Co. The 3154
 *Koppers, H. A. G. 3154
 Korenman I. M. 2938
 Korn, P. 2958
 Korsaw, N. 3054
 Kotowski, A. 2939
 Kotemana, L. 3109
 Kreck P. C. 2906
 Knechtbush J. G. 3143
 *Kreindlen G. 3170
 Kräncke G. 3174
 Krafz C. P. 3019
 *Kramer, E. 2832 3178
 Kramera, H. A. 2887
 Kraus, R. 3070
 Krause A. 2933
 Krause, H. 2925
 *Krhma I. J. 3013
 Krefa, O. 3187
 *Krebs & Co, G. m. b. H. 2882
 Krendler, A. 3069
 *von Kreuter, A. 3119
 Krenner, P. 3145
 *Kressman, F. W. 3183
 Kreslen D. J. W. 3148
 Knech, H. 3072
 Knezer E. 2937
 Kneha B. H. 2940
 Kneha S. 2981, 3001 3028
 Knehaewamy C. K. 2949
 *Kntchevsky, W. 3103, 3157
 Krivobok, V. N. 2952
 Krotova N. 3161
 Kruter, O. 2906
 Krueger, A. P. 3024
 *Krupp P. Gussion werk A. G. 2881
 Kruyt H. R. 3193
 Kuba B. 2943
 Kuchler, K. 2903
 Kueck P. D. 2923
 Kürscher, H. 2972
 Küster, H. 2922
 *Küttner, P. A. G. 3169
 *Kuhn H. A. 3100
 Kukolev G. V. 3142
 Kullerud G. 3126
 *Kulp, W. L. 3096
 Kunda, M. M. 3048 3061 3082
 *Kunt K. 3119
 Kuntz M. 2900
 *Kunsttänzer Patent Verwertung A. G. 3119
 Karoda C. 3011
 Kusnez A. 2953
 Kurth H. 3117
 Kurbayev M. A. 3053
 Kusmann A. 2957
 *Kutter A. L. 3178
 Kuten N. 3187
 Kutnetsov D. S. 3109
 Kurnetsov S. I. 3105
 van Laar J. J. 2901
 LaBarre, J. 3020
 LaBarre J. 3079
 Labat, R. 3075
 Labunsov, A. N. 2943
 Laevreng, P. 3018
 *Lafay, A. 2909
 *Lafite, T. 2883
 Laks, P. 3172
 Lakhani M. P. 2978
 Lakos P. 2956
 Lai K. 2901
 La Lande, W. A., Jr. 3162
 Lamar, J. E. 3113
 Lambert, A. G. 2951
 Lamiand H. 3024
 *Lamont, N. C. 3100
 Lamont, N. R. 3129
 Lancoz, A. 3078
 Landis, E. M. 3084
 Landsteiner K. 3056
 Lang, K. 3067
 *Langwell H. 2123
 Lanyi R. 2916
 Lanta J. 2133
 LaQue, F. L. 3165
 Lashov, L. T. 3072
 Larson, E. 2901
 Lachewer A. 2945
 Laudat, M. 3015
 Laurat, H. 3034
 Laurisch, A. 3026
 Lavetto, P. 3086
 *Lawrence, W. S. 3139
 Lawton W. E. 2849
 Lata, O. 3049
 Lazarev N. V. 3072 3074
 *Leach C. H. 3157
 Leck C. D. 3089
 Leeper, J. M. P. 2999
 *LeBaron P. E. 2882
 Lebedev A. A. 3173
 *Le cur hnt français 3196
 *Leck A. P. 2884
 Leemann W. G. 3105
 van Leeuwen, W. S. 3053
 Lefevre, M. 3030
 Legroux R. 3094
 *Leguillon C. W. 3199
 Lehmaon Pature H. 3058
 *Leibbraadt, P. 3119
 Leich J. N. 3039
 Leletz E. 3024
 Lemare A. 3077
 Lemke A. 2885
 Lemogne M. 3092
 Lendle R. 3076 3077
 *Leongyl B. 2801
 Leongyl V. 2893 2948
 Leonard G. P. 3127
 Leonard V. 3091
 Lemson, L. H. 3031
 Lepeshkin, V. V. 3060
 Lepper W. 3096
 *Lester N. 2974
 *Leventry, R. L. 2964
 Levi, G. R. 2915
 Levy, S. A. 3179
 Lewis, D. 2885
 Lewis W. C. M. 2896
 Ley, H. 2921.
 Libenson, L. 2909 3070
 Libera, F. 2955.
 Liddell E. G. T. 3083
 *Lide M. J. 2963
 Lieben F. 3001
 Liefink P. A. 3101
 van Lempt J. A. M. 2931
 Liesegang H. 3110
 *Liewer H. A. 3097
 Lightbody, H. D. 3044
 Lindemann H. 2916
 Ling A. R. 3120
 Ling S. M. 3037
 Linstrand, R. P. 2939
 Lipatov, S. M. 3161
 Lippmann, E. O. von. 3099
 Lipschitz W. 3076
 Lison, L. 3017.
 Lis G. 3090
 Lissarscey, J. F. 3129
 Lobel, L. 2929
 Lohmle, J. 3114
 Loch, L. 3066
 Loeper, M. 3077.
 Loevenhart A. S. 3089
 *Loew M. 3027
 Logan, L. 3151
 Lohmann, E. 3172
 Lohmle, J. 3098
 Loodfons J. 2941
 Long M. L. 3081 3089
 Lons - Ransow, P. 2960
 *Lorenz E. H. 3144
 Lower J. J. 3124
 Lovell, W. G. 3156
 Low J. A. 2994
 *Loyd R. W. 2963
 Lubbock I. 3141
 Lublin A. 3070
 Lubczyński O. 2911
 Lucary M. 3072.
 Lucentral R. 3187
 Lurhitzu V. I. 2947.
 Ludiny G. 3105.
 Ludwig H. 2909
 Lüthe H. 3121
 Lührke P. 2895
 Lunardoa L. F. 3023
 Lunde C. 2937
 Lundgren A. 2955
 *Lundgren, K. T. R. 2881
 Lurje M. 3054
 *Lutter, L. W. 3150
 Luetig, B. 3019 3022
 Luther R. 2879
 Lux H. 2935
 *Lyman, S. 2928
 Maass O. 2909, 3164
 MacArthur, C. G. 3127
 Macaulay, J. M. 2885

- *McAvoy, C. V. 3138
 McBain, J. W. 2906
 Macbeth, A. K. 3011
 McCarrison, R. 3031
 Maccoffa, G. 3064
 McClurg, H. B. 3035
 *McCullough, J. P. 2927
 *McDonald, J. U. 3154
 McDowell, R. J. S. 3047
 McGhee, J. L. 3036
 McGilivray, J. H. 3095
 *McGovern, W. J. 3138
 Macgregor, R. G. 3048
 MacGuffie, C. I. 2962
 Machatschki, F. 2912, 2943
 Maecht, D. I. 3078
 MacLunes, D. A. 2902
 MacKay, M. E., 3082
 *McKer, R. H. 3133
 *McKer, S. H. 2964
 McKennie, A., 2994
 Mackie, T. J. 3063
 *McKinley, C. W. 3160
 Macleod, J. J. R. 3047
 MacMahon, P. S. 2996, 2922
 *MacMillan, H. F. 3097
 McNallen, C., 3142
 *McNish A. F. 3141
 Macormac, A. R., 3173
 Macri, V. 3128
 *McWilliams A. C. 2883
 Maddock, S. J. 3039
 Madge, E. W. 2889
 *Madruck Ges für maschinelle Druck entwässerung m. b. H. 3152
 Magee, H. E. 3047
 Mahaut, P. C., 2918
 Majum, R. 2900
 Major, R. T. 2984
 Majumdar, K., 2915
 Majus, M. 3031
 Makarov, S. Z., 3132
 Makarova, K. A. 3016
 Mallé, B. 3109
 Malchevskaya, N. N. 3111
 Malfatti, J. 3041
 Malkin, T. 2967
 de Malleman R. 2887
 Mancke, R. 3075
 Mandelgrün, E. L. 3112
 Manjunath, B. L., 2981
 Mann, J. T. W. 2939
 Manske R. H. F. 3002
 *Manske R. H. P. 3131
 *Mantemach P. 2927
 *Manufactures de produits chimiques de Jouy - en - Joses (Anciens établissements Louis Des camps) 3141 3194
 Marbach K. 2913
 Marble, A. 3039
 *Marcard 3118
 Marcus, A., 3174
 *Marsznicz, C. 3147
 Marik, J. 3057
 *Margaria, R. 3047
 Mariller C. 3161
 *Marrasco G. 3069
 *Marrasco N. 2798
 Mark, R. 2950
 Marks, H. P. 3046, 3083
 Markson D. R. 3066
 *Marot, T. C. 3177
 Marotta D. 3123
 Marquardt, J. C., 3094
 Marshall, P. 3103
 Martin Sans, E. 3123
 *Martin Vialatte (Soc. A. r. b.), 3123
 Marx R. J. 3165
 Marzahn, W. 2961
 *Maschinenfabrik Benninger A. G. 3177
 *Maschinenfabrik Tilm. Gerber Söhne & Gebr. Wankleben. 3178
 Mass, F. 2929
 Massau, M. 2889
 Mason, J. H. 3064
 Mawrille, H. 3181
 Makhal, V. 2943
 Matrangoli, P. 2967, 3128
 *Masuda, S. 3178
 Mathers, F. C., 2961
 Mathieu, 3163
 Mathien, M. 2900
 Mathur, K. G. 2920
 Mathur, K. K., 2915
 Mathur, K. N., 2915
 Matignon, C., 3132
 *Matlock, C. 3138
 Matsubara, A., 2924
 Matzuda, T. 3061
 Matsui, M. 2901
 Matsuka, T. 3036
 Matthes, H., 2972
 Maucha, R. 3103
 *Mauler, C. 3137
 Maurice, B. 2944
 Mayerson, H. S. 3034
 Mayfield, C. R. 3056
 Mayr, E. 3094
 Mazzacova, G. 3080
 *Mechem E. R., 3129
 *Medinger, R., 3100
 Megalokonomos, J. G. 3187
 Mehta, S. M., 2902
 Meigh R. 3141
 *Meingard P. A. 3111
 Mellanby, M. 3039
 *Mellquist, J. H. 2927
 van Meis, W. H. 2841
 van Meisen, J. A., 2957
 Mendlik P. 3003
 *Menge A. C. 3104
 Menkin V. 3065
 *Mennie, F. L. 3152
 *Mennyei G. 3044
 *Menon A. S. 2495
 *Mensing C. E. 3016
 *Mentzel A. 3119
 *Merris M. H. 3132
 *Mers W. 3040
 *Mesick R. O. 2909
 *Meslin V. S. 2937
 *Messner A. C. 3042
 *Messner & Co. G. m. b. H. 2853
 *Metallgesellschaft A. G. 2924 2962 3131
 Mettler, S. R. 3091
 *Metz, Z. 3119
 *Meuser, L. 3199
 Meyer A. W. 2901
 Meyer B. S. 3031
 Meyer D. 3116
 Meyer, E., 2884
 Meyer, P. 2932
 Meyer, H. 2999
 Meyer, J. R. 3055
 Meyer, K., 3059
 Meyer, M., 2925
 Meyer, P. 3076
 *Meybaler P. 3064
 *Miebael L. 3038
 Miekley, I. J. 2911
 Miekley T. Jr. 2967
 Mier, W. 3112
 Miger, A., 3025
 *Mikhalovski, I. I., 2882
 Mikhailovski, A. G. 3116
 Mikumo, J. 3169, 3190
 Mibauer, J. 2929, 3181
 Milbourn, M. 3163
 Miles F. D. 3163
 Miles, G. H. 3111
 *Miller, R. 3153
 *Miller, F. B. 3158
 Milligan A. G. 2879
 Mills, C. A., 3042
 *Milne, S. 3170
 Min, P. 3090
 Minner, V. I. 2993
 Minetti, H. 2969
 Minot, G. R. 3091
 *Mirgala G. 3023
 *Missroh, A. 3089
 Mitchell, C. A., 2909
 Mitchell, J. A. 2972
 Mitchell, J. H. 2894
 Mitolo, M. 3039
 Mitra, S. K., 3028
 *Mitsun Kozan Ka-hushiki Kaisha, 3133, 3137
 Miyasaka M. 3173
 Miyata, M. 2929
 *Mizger, F. C. S., 2927
 Mlejnek, V. J. 3195
 *Moberg A. R. 3106
 *Möller & Meiser 2882
 *Möllerström J., 3064
 *Moserkofer W. 3031
 *Mogilucki R. 3090
 Mohlman P. W. 3104
 Moles, I., 2891
 Mohr, H. 3069
 *Moll, F. 3118
 *Moll H. 2881
 *Moll R. W. 2974
 *Möller, K. O. 3077, 3069
 Molnar F. 3091
 Moloney P. J. 3090
 Moncorps C. 3079, 3074
 *Moncrieff J. Ltd. 3111
 Monsacchi M. 3024
 Montagne P., 2883
 Montequi, P. 3021
 Montequi, R. 2934
 Montgomery, A. V. 3116
 Montonna, R. F. 3162
 Mooya, C., 2991
 Moos H. W., 3017
 *Moon J. W. 2966
 Moran F. 3037
 Morelli, J. 3061
 Morrell, M. 2900, 2990
 Morgan, B. G. R., 3035
 Morgan H. W., 3164
 *Morgan, W. H., Sr. 3101
 *Morrell, J. C. 3154
 Morrill, I., 3100
 *Mornell, E. 3137
 *Morrison, J. W., 3160
 Morse, C. W., 2935
 Moss, E. D., 2880
 *Mosso, G. 3177
 *Motters, H. H., 3095
 Mottuz, J. 3188
 *Moye B. W., 2963
 Müge, H., 3071
 Muehlberger, C. W. 3049
 *Mullensiefen, H. 3144
 Müller, D. 3031
 Müller, E. 2969
 Müller, H. 3152
 Müller, P. 3112
 *Müller, W. 2930
 Mukherjee, H. N., 3083
 *Mulholland, V. 3144
 Mullen, I. 3036
 Mullan C. E., 3163, 3173
 Munch, J. C., 3090
 Mundkur, B. B. 3029
 Munewada T. 3128
 Murphy, J. B. 3056

- Murray, G. P., 3193
 Murray, M. J., 2932
 Murty, K. S., 2921
 Muskhietov, D. I., 2946
 Mussong, P., 2942
 Muthanna, M. C., 2896
 Myhua, W., 3172

 *N. V. de Balaafche Petroleum Maatschappij, 3015
 *N. V. Electro-Chemische Industrie, 3134
 *N. V. Hollandsche Batunil Meatschappij, 3147
 *N. V. Mij Tot Beheer en Exploitatie van Oetronnen, 3144
 *N. V. Philips' Gloeilampenfabrieken, 2883
 *N. V. Sulca en Ovenbouw Mij, 3154
 *Nabreutellabrick Julius Frenner A.-G., 3131
 Naeer, G., 2911
 Nagai, L. J., 2915
 *Nagai, Y., 3131
 Nahocky, A., 3150
 *Nakahara, M., 3146
 Nakamura, K., 2971
 *Nakas, G., 3154
 Nakata, H., 2903
 Nakatsuchi, A., 3142
 Nakvins, D. V., 2947
 Nandi, B. K., 3002
 Narasimham, N. L., 2908
 Narayan, A. L., 2916
 Narayana, N., 3020
 Narayanamurti, D., 3029
 Nasuni, A., 2902
 Nasu, N., 2906
 Natarajan, C. V., 3054
 Natchou, S., 2983
 Nathan, L., 3121
 Natta, G., 2982
 Nathan Larmer, L., 3035
 *Naugatack Chemical Co., The, 3199, 3200
 Nava, V., 3028
 Navastro, I., 2907
 Navarro, L. F., 2945
 Neave, S. L., 3107
 Neall, J. M., 3054
 Nelson, E. E., 3091
 Nelson, E. K., 3095
 *Nelson, E. O., 2852
 Nelson, H. W., 2925
 Nelson, R. A., 2911
 Némec, A., 3110
 *Nendza, A., 2883
 *Neavler, F., 3118
 Nettmann, P., 3179
 Neuberg, C., 3023
 *Neumann, J., 3169
 Neuschloss, S. M., 3051
 Neville, H. A., 2927
 Newcomb, C., 3051
 Newall, I. L., 2937
 Newitt, D. M., 3099
 *Nichols, H. J., Jr., 3159
 Nichols, M. L., 2935
 Nicol, L., 3145
 Nieder, J. B., 2943
 Niethammer, A., 3030
 van Nieuwenburg, C. J., 2935
 Nieuwland, J. A., 2946
 Nikolaev, M. F., 3074
 Nikolaev, O., 3065
 *Niles, G. H., 3155
 *Nippon Kaisha Kabushiki Kaisha, 3119
 Nishimura, S., 3018
 Nishimura, Y., 2916
 Nishibori, R., 2904
 Nishida, K., 3042
 *Nishimura, T., 3131
 *Nishardy, P. W., 3130
 *Noble, O. C., 3144
 Noddaek, I., 2990
 Noddack, W., 2930
 *Noeggerath, J. E., 2927
 Norcom, G. D., 3101
 Normand, A. R., 2906
 2937
 Norma, J. H., 3046
 *Norika Aktiengesellschaft für Elektrotechnik Industri, 2963
 Northrop, F. S. C., 2909
 Northrop, J. H., 2900
 3024
 *Norton, P. H., 3144
 Norwick, I., 2970
 Nottus, P., 3092
 Nováček, R., 2943
 *Novák, H., 3160
 Nove, I., 3085
 Novoselsky, V. A., 3072
 Novotná, B., 2943
 Novak, S., 3064
 Nusser, W., 3009
 Nuccorini, R., 3027
 3028
 Nurse, C. J., 3105

 Oberg, S. A., 3048
 *Obermann, O. J., 3179
 Obiguo, A. E., 3070
 Ochao, K., 2916
 Odasky, N., 3094
 Oernkov, J., 3028
 Oesterlin, M., 3074
 *Oesterreichisch Amerikanische Slagstein A. G., 3153
 *Osterreichische Chem. Werke G. m. b. H., 2927
 O'Flaherty, F., 3125
 Ogg, R. A., Jr., 3001
 *Oghetti, B., 3136
 Oguri, S., 2901
 *Okkoti, M., 3135
 Ohlsson, R., 3018
 *Okada, K., 3119
 Okada, S., 3073
 Okamura, Z., 3155
 Okolov, P., 3052
 *Okuda, S., 3189
 Oláh, G., 3059
 Ollano, Z., 2919
 *Olson, P., 3167
 Olaso, A. R., 2922
 Oltsman, T. V., 3049
 Oriel, G. H., 3061
 Orlando, U., 3114
 Orlov, S. D., 3057
 *Orne, S. W., 2924
 *Ornfeldt, R. O., 3097
 Ornstien, I., 3021
 Ornstien, L. S., 2916
 *Orthner, L., 3136
 Orfner, G., 2891
 Osaka, K., 3119
 *Osaka Alkali Toshi Kabushiki Kaisha, 3119
 Owen, C. W., 2919
 Oshima, M., 3042
 Oslund, R. M., 3082
 Osa, J. P. W., 3099
 Osterhout, W. J. V., 3030
 *Osterreichische Chem. Werke, G. m. b. H., 2927
 Ostmeil, P., 2917
 *Ostro Research Laboratories, Inc., 3174
 Otsu, S., 3114
 Otsu, T., 3090
 Oswald, C. T., 2957
 Otani, B., 2958
 *Otani, T., 3133
 Otis, A., 2953, 2962
 Otis, R., 3060
 *Oyama, Y., 3135

 Pace, J., 3017
 Pachol, R., 3035
 *Pacz, A., 2966
 Padilla, T., 3087
 Padurova, O. G., 2945
 Pace, I. H., 3088
 Paez, H. A., 2939
 Pat, P. P., 2905, 3157
 Palacios, J., 2907
 *Palacky, P., 2928
 v. Palacky, J., 3059
 Pan, L. C., 2925
 Paneth, P., 2835
 Pancker, P. B., 2940
 *Panous, C., 3134
 Pape, A., 2839
 Papp, P., 2948
 Pappenheimer, A. M., Jr., 3048
 Paramawara, H., 2911
 Parange, G. R., 2903
 Park, C. R., 3197
 Parker, A. S., 3045
 Parmelee, C. W., 3143
 *Parnish, M. B., 3175
 Parshad, A., 2904
 *Parsons, C. E., 2965
 Parsons, H. T., 3036
 Passerelli, L., 2922
 Pasternak, L., 3055
 Pastorello, S., 2959
 Patel, J., 3077
 *Patent Facile & Color Corp., 3159
 *Patentverwertung A.-G. "Alpen" (Soc. Anon. pour l'exploitation de brevets "Alpen" Patente exploitation Cy "Alpen," Ltd.), 3131, 3136
 *Patterson, W. L., 2851
 *Paul, R., 3097
 Pauling, L., 2910
 Paulino, P. L., 3155
 Paathemer, 2993
 Pavolini, T., 2933
 Paz, P., 3099
 Payne, J. H., 2965
 Pearce, J. N., 2994
 Pearce, L., 3106
 *Pearson, T. F., 3146
 Peat, S., 3045
 *Pechin, L. C. F., 3137
 *Pechin, D., 2952
 Pehrson, E. W., 2950
 Pella, D. L., 3163
 Peller, J., 2902
 *Pellier, O. L. A., 3131
 Penayculik, S. W., 2946
 Penaval, G. H., 3089
 *Perron, M., 3184
 *Perry, R. S., 3185
 Peruta, A., 3019
 Pesce, B., 2933
 Peschke, W., 2941
 Pessel, L., 2960
 Peters, C., 3162
 *Peters, E., 3183
 Peters, K., 2972
 *Peters, O., 2851
 Petit Dutilleul, S., 3078
 Petras, M., 3151
 Petri, E., 3092
 *Petroleum Conversion Corp., 3158
 *Pew, A. E., Jr., 3159
 Pezzani, J. A., 3080
 Pfister, G. T., 3019
 Pfister, J. J., 3045
 *Pharmaceutische Werke "Morgine" A.-G., 3123
 Philipp, G. G., 3122
 Philipp, K., 2912
 Phillips, M., 2982
 Phukan, L., 3028
 Piatnicki, 3142
 Piatra, M., 2943
 Picon, M., 2994
 Pidgeon, L. M., 3136, 3164
 Piersanti, J., 3034
 *Pierce, R. H. H., 3144
 Pierce, W. O. D., 3179

- Neters, H. A. J., 3149, 3152
 Netsch E., 2939
 Nighal G., 3062
 Nix, L., 3092
 *Nix, R. D., 3133.
 *Nillings, N. B., 2965
 Nitschke M., 3038.
 Nussens, L., 3021, 3088.
 *Nitsch, J. A. G. 3153
 Piotrowski, W. J., 2967
 *Pipkin, W. A. 3100
 Pitombo, P., 3125
 Prowarsky, E., 2954
 Plack, M., 2903
 *Platsch M., 3152
 *Platt, H., 3177
 Plaut, P., 3054
 Plotnikov I., 2921
 Pogány, H., 2914
 Pohl, E., 2937
 Pokrowski, G. I. 3109
 Pollak, L., 3062
 *Pollock, R. T., 3158
 Polson, C. J., 3055
 *Polysius, G., A. G., 3146.
 Pomp, A., 2956
 *Pome, L., 2582
 Popp, 3093
 *Poppe K., 3054
 Popryadukin, A. A., 3195
 Porter, F. B., 3104.
 Portevin, A., 2956.
 Potek, D., 3020
 Potis, T. T., 3166
 Poulsen, K. A., 3026
 Pound, J. R., 2986
 Pourcelle, 3150
 Povoronskaya S. A., 3021.
 Powell, C. F., 2910
 *Power Gas Corp., Ltd., The, 3153
 *Powers, E. J., 3013
 de Prada, F., 3146
 Prasad, M., 2902.
 *Prett, W. B., 3167.
 Prusnitz, P. H., 2879
 *Preiswerk, E., 3015
 Prescott, R. T. M., 3118.
 Preston, F. W., 3141
 Preston, J. M., 3173
 Prevost, C., 2975.
 Price, W. A., 3035.
 Prikhodko, M. I., 3110
 *Pritchard, J. C., 3101.
 Pritsker, J., 3125
 Proske, H., 3105
 Probst, J., 2912
 Proverman, R., 3036
 Pruit, H., 3155.
 Prytz, M., 2905
 Przedniecka - Jedrzejowska A., 3024
 Prabram, K., 2890
 Ptaszek, L., 3071
 Publowl, H. E., 2956
 Puckner, W. A., 3130
 *Poemung, P. 2558
 Pogb, A. J., 3112
 Pugh, A., 3046
 Pulkayantha, R. M. 2922
 Purves, C. B. 3047
 Pusch, A. 2961
 Postumukov, N. N., 2951
 de Potnoky L., 2923 3132
 Quagliariello G. 3037
 Qureshi, M., 2921
 Raber, O. 3015.
 Raczkowski H. P. Z., 3109
 Radoff E. M. 3045
 Raff H. 3163
 Raffin R., 3064
 Raha, P. K. 2922
 *Raksanyi P. 3195
 Rakuzo M. A., 2932
 Ramage G. R., 3007
 Ramawamy, C., 2919
 Rambowek, F., 3192
 *Rambush N. E., 3153.
 Ramon G., 3064
 Ramsauer, C., 2910.
 Ramser, H., 2935.
 Rangel, O., 3064.
 Rao, R. S., 2924 2925
 Rao, G. G., 2904, 3029
 Rao, K. R. 2916
 Rao, L. R., 2949
 Rao, M. G., 2946
 Rao, M. G. S., 2945.
 Raper, R., 3006
 *Raschig F. G. m. b. H., 3012, 3153
 Rashevsky, N., 2904
 Rath, P., 3074
 *Rathbun J. P., 2923.
 Rathery, F., 3078
 Ratner, B., 3063.
 Randenbusch, H., 2969
 Raurieb, F. E., 2939
 Raux, J., 3120
 Ray, S., 2883, 2910, 2912, 2914, 2915.
 Rayleigh, 2918
 Raymond - Hamet, 3072
 Razumovskij, N. K., 2945
 Re, P. M., 3020
 *Read, F. F., 2923
 Rebeck, M., 2991
 *Rector, E., 3154
 *Reddelio, G., 2930
 Reding, R., 3053.
 Reed, L. J., 2905
 Reemak, P. H., 3034
 Reeser, O., 3070
 *Reeves, T. W., 3159
 Rege R. D., 3020
 Regendanz, P., 3064
 Rehberg, P. B., 3063
 Reichert, R., 2944
 Reid, E. E., 2972
 Rein, C., 2951
 *Reichardt, W. L., 3100
 Reiss, P., 3077
 *Reiter T., 3131
 Remond A., 3056
 Remy, P. 3072
 *Remy W. 2973
 Renaud M., 3023 3074
 *Rennerfelt I. 2963
 Rentz P. 3075
 Reppas A. O. 3124
 Rettiger L. P., 3076
 *Rewald B., 3195
 Rezer I., 2931 3141
 *Rheinhold & Co. Vermehrte Kieselguhr und Kockstein Ges., 3137
 Ribas I. 3146
 Rice P. O. 2967
 Rice O. K., 2911
 Richard L., 3033
 Richardson L. V. 3054
 Richardson O. W. 2915
 Richter, P., 3011
 Richter, G. A., 3164, 3169
 Richter, K., 3094
 Riddle, P. H., 3143
 Ridgway, J. H., 3179
 *Ried, O., 3130.
 *Riedel, J. D. F. de Hato, A. G., 3014
 Riedel, W., 3122
 *Riedhammer, L., 2953.
 *Riemann, O., 3194
 Riet, E. D. 3150
 *Ries, H., 3178
 Ries W., 2884.
 Riewer O., 3069
 van der Riet Woolley, R., 2918
 Rieger, G., 3102
 Rinne P., 3141
 Risch, D. McK., 3043, 3053
 Rapperton J. C., 3093.
 *Ras, K. B., 3154
 Ritchie, P. D., 2994
 Ren, J. O., 3125
 Robbins, F. M., 2954
 Robersht, A., 3049
 Robertson, W. C., 3114
 Robinson, M. E., 3079
 Robinson, P. L., 2933
 Robinson, R., 3004
 Robinson, R. A., 2903
 Robison, R., 3040
 Roeben, M., 3031.
 Röders P., 2584
 *Römer A. Flärberei Bleichen und Appreturanstalt, 3176.
 Römler, H., 3114.
 Röttger, G., 3096.
 Roda, A. H., 3056
 *Roger, M. S., 2964
 Rogers, J. A., 3066
 Rohmer, P., 3039
 *Rohn, W., 2924
 *Rottbeim, A., 2963
 Roldán, J. C., 2935.
 Ro"er, F. S., 2905
 *Roman, R. P. L., 2943
 Roman, W. 3021, 3044
 Romano, S., 3063
 Romanova K. 3099
 Romero G. 2915.
 Romoli Venturi D. 3046
 Romwalter A. 2913
 Roncato A., 2979
 Ronchene, P. 2919
 Rondou, P., 3041, 3044
 Roos, O. 3112
 *Rosenblatt, A., 3054
 *Rower, I., 3153
 Koska G. 3080
 Roth P. 3035
 *Rothers, W. S., 3185
 Rothemann, M., 3051
 Rothmann H., 3070
 Rovank A. P. 2912
 Rowe, P. W., 2953
 *Royer, E., 2927
 Roy, M. K., 2946.
 Roy, S. C., 2976
 Roy, S. K., 2945.
 de la Roza, J., 3161
 Rozanov, A. N., 3111.
 *Ruben, S., 2926
 de Rubles S., 3104
 Rutano, P., 3063.
 Rubino - Montevideo M. C., 3064
 Rubner, M., 3037
 *Rubby, G. H. A., 3149
 Ruchbolf, C. C., 3103
 *Rue, J. D., 3169
 Rüb, J., 3093
 Rüdiger, 3120
 *Rüberg, F., 3134
 Ruiz, C., 2999, 3079.
 Rule, A. M., 3066.
 Rupe, H., 2990
 Ruppel, A., 2991.
 Ruppel, E. 2922
 Rushton, W. A. H., 3017.
 *Russ, A. J., 3100
 Ruzicka, L., 2957, 2958
 *Ruzicka, L., 3012.
 de Sábada R., 2938
 *Sabatseha, T., 3131.
 *Sacharin Fabrik A. G. vorm. Fallberg, List & Co., 3190
 Sacher, J. P., 2924
 Sachs, G., 2957, 2958.
 Sachs, H., 3059
 Sachs, L. M. G., 3047.
 *Sadu, M., 3190
 Sagastume, C. A., 3080
 Sager, O., 3055
 Sab, P. P. T., 2976
 Saha, H., 2933
 Saha, M. N., 2910, 2920
 Saida, T., 3188.

- Sack S 3062 3063
 Sallard 3194
 Saward L 3031
 *Sauderschna N 3163
 Sauto K 3060
 *Sakata T 3115
 *Sala C J 3176
 Saladin R 3151
 Salatore S A 2999
 *Salisbury H D 2923
 Salmoraria E 3155
 Salmon E S 3122
 Salmon - Legagneur F 2990
 Salomon A 2949
 Salomon R 3076
 Samhamurti G 2916
 *Samburov S V 2941
 Sambursky S 2916
 Sanders K 3191
 *van de Sandt H 3123
 *Sanger W E 3191
 Santenouze D 3079
 Sapozhnikov A 3171
 Sapper A 2945
 Saradzichiv P 3064
 Sare J 2942
 Sarjant R J 3143
 Sarkar A N 2993
 Sarus H R 2913
 Sartori A 3089
 Sasaki A 2929
 Sasaki N 2921
 Sasaki R 3036
 Sasaki T 3065
 Sattri H N 3025
 *Satchwell L 2983
 *Sato H 3176
 Sato M 3017
 *Sawot T 3190
 Sauerwald F 2957
 Sawada T 302
 *Sayers W W 3108
 Scharne B 2937
 Scharrer K 3117
 Schay G 2985 2993
 Schaefer P E C 2909
 *Schaber J 3164
 Schena A 3069
 Scheps M 3071
 Scheninger W 3041
 *Schering Kahlbaum A-G 3012 3014 3121
 Scherner S 3053
 Scherp H W 2991
 Schilling V 3072
 *Schiller W 3024
 Schlenker H I 2931
 Schlesmans C 3060
 *Schlitt J L 2941
 *Schlögl R 3119
 Schlossmann K 3042
 Schmid R 3070
 Schmid W 3103 317
 *Schmudding W 3165
 Schmidt A A 3172
 Schmidt D 2941
 Schmidt H 3163
 *Schmidt J H 3136
 Schmidt L 2961
 Schmidt Nielsen S 3161
 Schmitt F O 3044
 Schmitt L 3114
 *Schmitt H P 3012
 *Schmitt T 3125
 Schofield R K 2999
 *Schödlauer A 3145
 *Schmoka H 3170
 Schaeffer K 3155
 Schneider J 2944 3037
 Schneider K 2992 2993
 Schneider W 2946
 Schaefferbach H 2934
 *Schneppensfabrik Frankenthal & Cie A-G 2974
 *Schattpahe K 3012
 *Schöller-Bleckmann Stahlwerke A-G 2965
 *Schöller W 3012
 Schorn M 3044
 Schorn M J 2945
 Schorn R 3065
 Schönbeimer R 3051
 Schoep A 2941 2945
 Schofield R K 2999
 Scholder R 2935
 Scholl R 2996
 Schottan P 2940
 Schurr T 3047
 Schroeder E P 3009
 Schrüfer P 2911
 *Schroeter G 3015
 Schroeter G 3030
 Schubert J 3034
 Schults P W 3082
 Schults G 3174
 Schulze B 3164 3166
 Schulze K 3127
 *Schupp H 3166
 Schuske W 2934
 Schuster K 3124
 Schuster P 3151
 Schwartz L H 3045
 Schwartz M 2908
 Schwartz M A 2937
 Schwarz H 3072
 *Schwarzkopf H 3176
 Schwetzer F 2973
 Schwenker P 2941
 Scott A W 3094
 Scott C M 3089
 *Scott W C 2963
 *Scottish Dyes Ltd 3170
 Scott G 3037
 Searle A B 3142
 *Sebold & Neff 2964
 Sebrill W H 3035
 *Seef 2925 3077
 *Segal L 2926
 *Seibert J A 3190
 Seidenberg S 3049 3169
 *Seidenschaur F 3169
 Sekama J 2944
 Selden H H 3127
 Seligmann M 3132
 *Sellet L 3186
 Selzer G E 3023
 Semb J 2955
 Sembratka W 3165
 Semenov V V 3123
 Semp H 2996
 Semova O M 3057
 Sen B M 2912
 Sen H K 2988 2995 3157
 Sen N R 2988
 Sen R N 2991
 Seogoku T 2999
 Serdyuchenko D P 2944
 Sereni F 3062 3067
 Serpyrka M V 3077
 *Serzhenko P 2990
 Serra R 2923
 *Serris R 2980
 *Seyffert E 2982
 Seyler C A 3149 3151
 *Shaliova P N 3147
 Shah M S 2907
 Shakhbeldise A B 3195
 *Shakino A P 3148
 Sharma S S 2916
 Sharp B R D 3174
 Shaw L A 3042
 *Sheldon S B 2964
 Shepard A P 2967
 Shephard H H 2952
 Sherman H E 3173
 Shuman S 2901
 Shumayama K 3181
 Shunoda S 2944
 Shuraw G 3149
 Shurata M 3073
 Shushko H 3002
 Shushkova S 3060
 Shve J W 3033
 Shvuk A 3124 3125
 Shvuk A A 3124
 *Shvovoch H P 3139
 Shvire R N 3171
 Shur M P 3139
 Shuya L 2916
 *Shuya T 3136
 Sibon G 2954
 Siebert K 3120
 Siskarenko A L 3192
 Siebert W J 3081
 Sievert F 3040
 *Siemens Schuckertwerke A-G 2923 3177
 *Siemens F A-G 2964
 Silva L 2999 3079
 Silvestrovich I I 3097
 Sim A H 3033
 Simakov V N 3111
 Simon A 3040
 Simon E 3023
 Simon J 3192
 Simond A E 3042
 Simons J H 2987
 Simonson J L 2986
 Simmons W H 3125 3189
 Singh B K 2942
 *Skrachenko I A 3116
 *Skarynski B 3047
 Skorchaya D 2912
 Skola V 3142
 Slater E C 3153
 Slavik F 2944
 Slocum M A 3044
 Slove A 3043
 Soria E H 2993
 Smurov N D 3113
 Smut A J H 2946
 Smith B A 3174
 Smith E 3060
 Smith G 2941
 *Smith G P 2981
 Smith G V S 2982
 Smith L A 3133
 Smith P W 3043
 Smith R S 3081
 Smithells C J 2923
 Smoker E H 2901
 Smolczek E 3095
 Smolidski K 3192
 Smolik L 3109
 Smorodintsev I A 2948
 Smyth C P 2947
 Smyth H F 2996
 Sotkarev A N 2947
 Soutka K 3171
 *Soc. anon. de Commeny Fourchambault & Decaenille 2965
 *Soc. anon. de decoration architecturale et mobilière 3137
 *Soc. anon. des établissements Tégalemont 2951
 *Soc. anon. pour l'Ind. chim. & Bâle 2926 3012 3175 3176 3177 3200
 *Soc. anon. pour l'Ind. de la lagouste (Succursale de Hougne) 3147
 *Soc. anon. des manufactures des glaces et produits chimiques de St. Gobain 3144
 *Soc. anon. le nickel 2983
 *Soc. anon. Frétilles et laminoirs de Havre (Anciens établissements Lazart Weiler 3146
 *Soc. chim. de la Grande Paroisse (A. zote et produits chimiques) 3119
 *Soc. coopérative de Rugles et lac analis non électrique réunis 3146
 *Soc. electrics ed. el. trochimica del Calaro 3134

- [illegible]

- Uchono M 3181
 Ueda H 3055.
 Ueda S 3189
 Ueno S 3187, 3189, 3189
 *Uick A 3123
 *Ukrainskoe Aktsion-
 ernoe Obshchestvo
 po palniam i rali
 zaluz isobreteni
 Lakons 2881
 Underhill P P, 3043
 3041
 Underwood H W.
 Jr 2995
 *Unger M 2928
 Urbanek L 2937
 3035
 Uaso V V 2951
 Uspenskaya Z P,
 3114
 Utkina, Lyubovtsova
 K., 3072
 Vácha J, 3101
 Valdevelleno C 2910
 Vallagnone L, 3073
 Valley Radol P
 3055
 Vallée B, 3101
 Valis B H, 3150
 Vandaveen, P G,
 3151
 Van Dyke H B, 3099
 Van Heuckeroth, A.
 W, 3132
 Van Kirk R W, 2982
 *Van Nuy, C C.,
 2891
 Vanderwall J W.
 2946
 Vandewelde, A J J.
 3111
 Van Valkenburgh H
 B, 2931
 Varma P S, 2980
 2981, 2983, 2994
 Varma, S C., 2896
 2909
 Vaskko, J, 3191
 *Vaseua, A E, 3191
 Vassovitch N B,
 2945
 Vastagh G, 2939
 *Vaughan, A H, 2964
 vander Veen H., 2972
 Veithmeyer F J, 2979
 Veil S, 2899
 Veluz L, 2905 3021
 Ven H A 2949
 Vendl M 2912 2913
 Venkatesachar B,
 2918 2925
 Venuto L J 3133
 Verbeelen A 3111
 *Verro für chemische
 und metallurgische
 Produktion 3015
 *Vereingte chem Fab
 Kendl Heller & Co
 3115
 *Vereingte Cigh
 lampen und Electr
 cills A-G 2883
 *Vereingte Silberham
 merwerke Hertzel
 & Co, 2965
 *Vereingte Stahlwerke
 A-G 2964
 Verboek P H, 2904
 Verster, P, 3031 3097
 Vessilka J, 2961
 *Veyron P M, 3170
 Viale G 3063
 Vland P 3097
 Vickery J R 3187
 Vidacovitch M, 3079
 Viktorov, P P, 3173
 Vila R C., 2942
 Vinetskaya E Y.,
 3181
 Virtanen A I 3030
 Vischer, M B 3043
 *Vistarial M, 3194
 *Vitas P S, 3170,
 3185
 Via Limuro, G D 3018
 Vlassopoulos V, 3011
 Vles, P., 3064
 *Völker, E, 2953
 *Vorel Jorgensen M,
 3146
 Vogt R, 3012
 *Vogtlander H, 2966
 Volk, H, 2984
 Vollmer, H, 3070
 Volmer V, 3121
 Volz, J L, 2971
 Vondrák, J., 3191,
 3193,
 Voroshilov, N N,
 2902
 Vyzvolil, J, 2944
 von Wachenfeldt R,
 2946
 Wachwans M M,
 2945
 Wadsworth J M,
 3135
 *Wagner, P H, 2982
 *Wagner, G (Firma)
 3170
 Wagner H, 2992,
 3181
 Wagner, O H, 2933,
 Wagner W 3025
 Wahn B 3087
 *Wast J P, 3099
 Wakeman R L 2995
 Wakman S A 3119
 Walawski J 3098
 Walbum L B 3066
 Waldschmidt Lesa
 E 3019
 Walker P 2971 3128
 Walker M A 3082
 Walker T K 3122
 Walland H 3174
 *Wallis J S 3159
 *Walker G 3147
 3183
 Walz K 3146
 Wan S 3037 3038,
 Wang C C 3060
 Warner A E 3106
 *Warren A T W,
 3135
 Warren R G, 3112
 Warren, W J A 3110
 *Wadno, Y, 3177,
 Waterman H I, 3187
 *Watkins, G D, 3144
 Watkins O, 3042
 Watson A H., 3095
 Watson P D 2995
 Watt, J M, 3127,
 Welch, K, 2921
 Webster, R W 3019
 Weed, A M 3043
 Weiss, H, 3075
 Wehmer, P, P, 3166
 Wehrle P, 2906
 *Weidach, W, A.,
 3145,
 Weil, A, 3069
 Weid H, 3043
 *Weiland, H J, 3173
 Weisberg S J, 3040
 *Weisland R 3165
 Weser, H B, 2897
 *Weichen, C, 3170
 *Wesse J M, 3135
 Wesskopf V, 2907
 Wesel W, 2917
 Weller, W. C., 3131
 *Weller, A, 3191
 Wendler A, 3141
 Wenger P, 2934
 West, I, 3020
 West, I S 3054
 Wertheimowas, M.,
 3192
 Wesley W A, 3135
 West C J., 2909
 West, J H, 3035
 Wester, D H, 2894
 Weston R S, 3101
 Westwood J B, 3121
 deWet J P, 3160
 Weyer P, 2955 2962
 *Whelan J J, 2926
 Wheeler T S, 2890
 White A N 3130
 *White, C. O, 3189
 *White G H 3100
 White M 3128
 Whitehead T H 2997
 *Whitehead W, 3169
 *Whitfield C, 3154
 Whittemore L C
 3106
 Wibaut N L (Mme.)
 3105
 Wicker S A 3193
 *Wickop L 2983
 *Wiegand P 3200
 Wierl R 2896
 Wiertelak I 3163
 Wiese J 3072
 *Wietzel R 3180
 Wigley C G 3107
 Wilder R L 3052
 *Wildermuth W,
 3144
 Wilhelm C J 2961
 Wilkins P J, 2905
 Williams, C F, 2950
 Williams, C L 3118
 Williams, W P 2918,
 *Williams W M,
 3136
 Williamson J., 3141
 Williamsou, R. V.,
 2995,
 Willbiller, R, 3097
 Wilson, H L, 3094
 Wilson, W. K., 3094
 Wind, A. H., 2933
 *Windsheim, H.,
 3123,
 de Winwarter, H.,
 3035,
 Winkelmann H.,
 2980
 Winkelmann, H A.,
 3196
 *Winkelmöller E &
 Co, 3097
 *Winkler C, 3146
 Winkler J, 2879 2967
 Winnacher, K, 2879
 *Wintow C. A., 2881
 Winterfeld, K., 3005,
 3006
 *Wintermole, H A.,
 2928,
 *Wintershall, A-G,
 3151
 Wirth, C, 3121
 Wirsbaky, E, 3058
 Wodon J L, 3079
 Wölmer, W 3122,
 Woll K., 2910
 Woll K. L, 2921
 *Wolff et Co Kom
 mssid-Ges. auf
 Akure, 3168
 Wolfenstein W, 3072
 Womack, E B, 3044
 Wood, J C., 3030
 Wood, W. A, 2959
 *Woodall - Duckham
 (1920), Ltd, 2883
 3183
 Woodhead J E.,
 3120
 Woodruff S, 3006
 Woodward L A.,
 2918 2919
 Wrede P, 3040
 Wrochl G F, 2997
 Wright G P., 3050
 Wu H 3037
 Wünnenberg E, 2885
 Wulf P 2900
 Wuxter K., 3094
 *Wuxter, O H, 3190
 Wuxter, W, 2935,
 van Wyk A, 3034
 Yagata, H, 3021, 3084
 Yaginuma, T., 2973
 Yakimach A, 2935,
 Yamamoto, T, 2907
 *Yamanous G, 3178,
 Yang C, 3086
 Yanovski, V, 3193,
 Yap C, 2907,
 Varusoe S, 9 3110
 *Vales, W. H., 2891
 Vermolyeva Z, V
 3032
 Yokota K, 3063
 Yoshimura K., 3042
 Yoshitomi, M., 3058.

- *Youngman, R. H., 3144
 Yakimori, T., 3184
 Yar'ev, I. K., 2977
 Zachariasen W. H., 2942
 Zagami V., 3088
 Zahn H., 2910
 *Zahn & Co. Bau chemische Fabriken G m b H., 2963
 Zabo, O., 2580
 Zak, E., 3076
 Zaleski, J., 3193
 Zanotta V., 3124
 Zastrop H., 3052
 Zapolski V., 3124
 Zapp, E. V., 3090
 Zaprometov B. G., 3109
 Zarni, P., 3092
 Zaucher E., 3068
 Zavaritski, A. N., 2942
 Zechmeister L., 3029
 *Zeh R., 3133
 Zewa, H., 3072
 Zehkman I. F., 3192
 Zelnuski, N. D., 2979, 2977
 Ze"ner, J., 3032
 *Zemlyanitsin N. A., 2982
 Zhukovskii M. D., 3164
 Ziebeck, K. V., 2844
 Ziemer P. H., 3034
 Zih A., 3047
 Zikmunds, R., 3032
 *Zlotenski, I. G., 2982
 Zol'ner, E. A., 2979
 de Zombory, L., 2939
 Zubov V., 2942 2943
 Zubis M., 3053 3054
 Zumstein, O., 2974
 Zunt E., 3079
 *Zurlich A. J., 2966
 Zwickler J. J. L., 3126

CHEMICAL ABSTRACTS

Vol. 25.

APRIL 10, 1931

No. 7

1—APPARATUS AND PLANT EQUIPMENT

W. L. BADGER

A new high-efficiency glass condenser. PAUL R. CRANDALL *J. Lab. Clin. Med.* 16, 89-91(1930).—A glass condenser is described which may be used in an all glass distg. system. Approx. 2 l of distd. H_2O per hr may be produced. F. R. MAIN

A simplified apparatus for the extraction of liquids. R. B. FORSTER *Chemistry and Industry* 50, 44(1931).—The app. was designed for extg. H_2O solns. with lighter solvents. A tube from the closed bottom of the cylindrical extractor connects with the boiling flask. A by-pass from this tube enters the extractor about half-way up, serving as a passage for the vapor to the condenser on top and as a return for the solvent to the boiling flask. The soln. to be extd. is placed in the lower half of the extractor, the condensate falls into a thistle tube which delivers it at the bottom of the soln., through which it rises and overflows to the boiling flask, its rise being impeded by baffles attached to the stem of the thistle tube. *Cl. C. A.* 20, 3817. J. H. MOORE

A separatory funnel for washing heavy volatile liquids. HAROLD S. KING *Proc. Trans. Nova Scotian Inst. Sci.* 17, 240-1(1930).—Two separatory funnels are sealed together in such a way (see fig.) that a heavy liquid in the first can be sepd. from a supernatant liquid and delivered to the 2nd funnel without contact with the outside atm. After removal of the supernatant liquid, the heavy liquid can be returned to the first funnel for rewashing. RACHEL BROWN

An improved pipet manipulator. D. C. B. DUFF. *J. Lab. Clin. Med.* 15, 1027-8(1930).

Improved slow-combustion pipet for gas analysis. H. R. AMBLER. *J. Sci. Instruments* 8, 18-9(1931).—O is led into the gas through a capillary side tube entering near the bottom of the Hg in the pipet, thus preventing back-diffusion of the gas. Analyses by absorption in liquids may be made by bubbling the gas through the side tube. J. H. MOORE

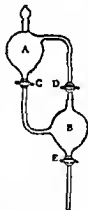
Automatic gas-analysis apparatus based on the thermal conductivity of gases, and its use in industry. PAUL JARRIER. *Chaleur et Ind.* 11, 181-5(1930).—A table of densities, viscosities and thermal conductivities of air, N_2 , O_2 , CO , CO_2 , water vapor, H_2 , C_2H_2 , CH_4 , lighting gas, NH_3 , and SO_2 is given. The Ranarex app. is described, it gives good results. The principle of the Dommer "Unigraph" is illustrated and described. The Siemens elec. CO_2 indicator is described. The elec. measurement of the thermal conductivities of mixts. of gases and its applications are discussed. S. L. B. ETHERTON

A thermal method for the rapid measurement of carbon monoxide. GERHARD STAMPE. *Messtech* 6, 267-70(1930).—The Dräger app. is described. It depends on measuring the temp. of the gas on entering and leaving a catalyzing chamber in which it is oxidized to CO_2 . The theory is discussed, with 7 references. *Cl. C. A.* 24, 5897. J. H. MOORE

Determination of humidity. G. ALIVERTI. *Nuovo cimento* 6, cxvii-cxxiii(1929).—A discussion of the use of various types of hygrometer. B. C. A.

A convenient hydrometer for determining the specific gravity of heavy liquids. A. C. TESTER. *Science* 73, 130-1(1931).—The app. was devised for liquids of sp. gr. between 2.0 and 5.0 for use in the sepn. of minerals. Only 5 cc. of the liquid is necessary. The sp. gr. of solids may also be detd. J. H. MOORE

An improved portable pH apparatus. ARAO ITANO. *Ber. Ohara Inst. Indus. Forsch. (Japan)* 4, 471-4(1930), *Cl. C. A.* 23, 3602.—The app. previously described has been changed to make one instrument serve as both galvanometer and millivoltmeter; it is manipulated by means of a newly constructed double-throw switch. The



wiring diagram and a photograph of the app are shown. The instrument measures the e m f of the cell quinhydrone, standard Pu soln | saturated KCl | sample, quinhydrone. N. A. LANGE

Deposition filter process. ARMIN BILMERTHAL. *Pharm. Presse* *aus prakt. Hefi* 1930, 165-6—Several illustrations are given of an improved type of filter for large operations, it has compartments filled with frames carrying wire gauze on which the filter mass (asbestos or other suitable material) is flowed and deposited in thin paper like layers. The various operations are carried out under pressure and suction. W. O. I.

The Schacht-Wheel drier. VON BEZOLD. *Chem. App.* 17, 217-8(1930)—The drier contains a series of cylinders vertically arranged, over which the material is passed and drops to the wheel below. Hot air is blown upward. The drier will not operate with materials which are larger than hazel nut size or which are sticky or liable to cause explosion. M. C. ROGERS

The sucking off and precipitation of dust and vapors. A. VOGT. *Chem. App.* 17, 209-10(1930)—The app. is arranged so that gases contg. dust are put in contact with a stream of water and passed into a settling tank. The sludge settles and the clear liquor overflows to a launder and is pumped again to the mixing chamber. The settled sludge can be pumped off through a discharge opening in the bottom of the settler. M. C. ROGERS

An instrument to measure soil consolidation. CORNELIUS DAVIES. *J. South-Eastern Agr. Coll., Wye, Kent*, No 27, 225-7(1930)—An instrument for measuring the compactness of soils consists of 2 telescopic tubes within which is a supported spring balance so arranged that a push on the handle of the appliance extends the spring, a scale calibrated in 0.5 lb., a max. reading indicator; and a set of steel probes of different lengths and diam. K. D. JACOB

Illumination of the half-shadow polarimeter with a mercury-quartz lamp. HANS N. NAUMANN. *Biochem. Z.* 229, 269-70(1930)—The use of the Hg-quartz lamp as a source of illumination increases the accuracy of the polarimeter 2-3 times, this is of great significance in the study of weakly rotating substances. S. MOROGLIS

Theory of optical pyrometers and description of some new improvements. S. HILD. *Chaleur et ind.* 11, 403-10, 473-80(1930). S. L. B. EMMETT

Leonard Hill acetone-methylene blue actinometer. T. W. B. OSBORN AND A. D. STAMMERS. *Brit. J. Actinotherapy* 5, 68-9(1930)—Exclusion of air is essential for securing reproducible results. B. C. A.

An air-thermostat for corrosion research. U. R. EVANS. *Chemistry and Industry* 50, 66-7(1931)—A uniform temp. is maintained by forced circulation of air from the chamber through channels between the double walls and back to the chamber. The elec. heating is controlled by an adjustable contact thermometer. J. H. MOORE

Photoelectric cell has definite place in color measurement (SCOTT) 25. Electrometric measurement of very low pressures by means of ionization currents, the space charge manometer (MOLTHAN) 2. Apparatus for determining the tendency of coal and other materials for spontaneous combustion (KREULEN) 21.

DUBSKÝ, J. V. Selbsttätige Filtrationsapparate. Leipzig Georg Thieme 64 pp. M. 3. Reviewed in *Chem. News* 142, 110(1931)

RABALD, ERICH. Werkstoffe Physikalische Eigenschaften und Korrosion. Band I. Allgemeiner Teil. Metallische Werkstoffe. Band II. Nichtmetallische Werkstoffe. Leipzig Reudnitz Verlag von Otto Spamer. Stitched, M. 128, bound, M. 135

Gas filters. RENÉ LAHAUSSOIS. Fr 693 035 Mar 29, 1930. Filous surfaces are obtained for gas filters with filaments submitted to a treatment such as the action of an electrostatic field so that they become stuck automatically at one end to an adhesive previously applied to the filtering surfaces.

Rotary screen and filter apparatus for reclaiming solid particles from gases. FRANK J. McDEVITT. U. S. 1791,102, Feb. 3. Structural features.

Centrifugal dust separators. Soc. HENRY SIMON, LTD., and GEORGE WATTS. Fr 693,823, April 12, 1930.

Combined hydrometer, liquid casing and suction bulb. J. H. COLLIE. Brit. 335 698, Aug. 29, 1929.

Humidistat. EDWARD A. POSTER. U. S. 1,791,375, Feb. 3.

Calibrated viscometer. RICHARD VON DALLWITZ-WEGNER Ger. 515,279, Jan. 19, 1926

Hydrometer syringe for indicating specific gravity of liquids. FELLIS H BRIDGE (to Electric Storage Battery Co.) U S 1,700,696, Feb 3 Structural details

Apparatus for the agitation and sedimentation of liquids and solids in suspension. THE DORR CO Fr 692,707, Mar 26, 1930

Apparatus for delivering liquids in measured quantities. WERNER HANDELSGES Ger 515,278, June 4, 1927

Apparatus for evaporating and concentrating liquids by passing them through a heated centrifuge MASTRIS J KRAMER U S 1,701,317 Feb 3 Structural features

Apparatus for effecting hydrogenations and other catalytic reactions between gases and liquids. RING GEA CHEM UNTERPHILINGEN M N H Ger 515,417, Jan 25 1922

Apparatus for treating gases with atomized liquids. STILLWATER A G VORN WISCH & Co Ger 514,413, Apr 3 1929 Addn to 512,128 (C A 25, 850)

Apparatus for regulating and indicating the density of mixed gases. SMOOT ENGINEERING CORP Ger 515,257 Jan 29 1928 See U S 1,682,689 (C A 22, 3409)

Gas washing apparatus. H A BRASSPAT (to H A Brassert & Co., Ltd.) Brit 335,618, July 6, 1929 Various structural details are described of an app comprising alternate rows of oppositely rotating or fixed and rotating bars grouped around an axis and fixed nozzles spraying the bars in an outward direction Cf C A 24, 1274

Apparatus with shields for reducing radiation effects in gas temperature determinations. ROBERT F WILSON U S 1,701,020, Feb 3 Structural features

Surface condenser suitable for condensing steam. CHARLES A PARSONS and FREDERICK ASKLESS (Askless to Parsons) U S 1,700,831, Feb 3

Evaporator calandria, etc. OSCAR H WURSTER U. S. 1,701,296, Feb 3 Various details are described of an evaporator suitable for evapn of crude glycerol

Shaft drier, with grinding or comminuting means. BERNHARD SAGENANTH Ger. 515,443, Oct 21, 1927

Centrifugal apparatus for dewatering lintera, wood cellulosa, etc. GYNS BILLMER MASCHINENFABRIK Ger 515,463, Nov 22, 1929 Addn to 499,079 (C A 24, 3139).

Röntgen-ray tube. C. H. P. MOLLER A-G Brit 335,510, Aug 27, 1929 Structural and elec. features

Cathodes for electron tubes. HERMANN LOOSLI Ger 515,095, Jan 25, 1924. Cathodes, comprising an alk. earth oxide on a core of difficultly fusible metal, are prep'd by dipping the core into molten alk. earth metal, which is heated *in vacuo* or in an inert atm., and then oxidizing the metal coating on the core Alternatively, the core may be embedded in a powd. mass of alk. earth metal which is heated *in vacuo* or in an inert atm.

Electric discharge tubes. J. E. LILIENFELD. Brit 335,537, Oct. 3, 1928 Various details of manuf. are given of tubes which may be used for the production of Röntgen rays and which may have a cathode of W or Ta pointed by etching and activated by Cs evap'd in an adjacent tube.

Photoelectric cell. N V PHILIPS' GLOELAMPENFABRIEKEN, Fr 693,013, Mar 29, 1930 A photoelec. cell contains a photoactive substance such as Cs spread on an intermediate layer contg. a chem. comp'd and particles conducting electricity. The Cs is produced in the cell or in a chamber communicating with it by heating a mixt. contg. a comp'd. of Cs such as Cs chromate and a reducing agent such as Zr. The mixt. also contains a chem. comp'd. which during the heating is volatilized or reacts with the reducing agent present to form the intermediate layer Cf C. A. 24, 2330.

Muffle furnace for calcining "chemicals," pigments or other pulverulent materials. R KRAUSSE Brit. 335,659, July 11, 1929 A muffle or retort extends through a furnace fired by burners and is suspended by blade springs so that it may be continuously reciprocated to convey the charge gradually from the inlet to the outlet end Various details of construction are described Cf C A. 24, 3106

Apparatus for controlling furnace combustion in accord with furnace boiler pressure variations. CHARLES M. TERAY (to A W Cash Co.) U. S. 1,700,685, Feb 3

Retort for stoker furnaces. WM. F. RIEGER (to Stoker-Matic Corp.) U S 1,790,840, Feb 3

Smoke inspection apparatus for boiler furnaces. ROBERT H. WAGER. Ger. 515,258, Feb 24, 1929 Corresponds to Brit 327,812

Roofs for industrial furnaces. MAURICE GILBERT. Fr. 693,588, April 9, 1930

Suspended roofs for furnaces. SOC. ANON. DES POYERS AUTOMATIQUES Fr. 692,542, Mar. 21, 1930

- Closures for annealing or melting furnaces. AKT.-GES. BROWN, BOVERI & CIE Fr 693,214, April 3, 1930
- Feeding device for crude lignite trough-grate furnace. FRANKEL & VIEHANN Ger 515,403, Dec 8, 1928
- Tuyère for underfeed stokers. ROBERT A. FORESMAN and DONALD J. MOSSHART (to Westinghouse Elec and Mfg Co.) U S 1,791,269, Feb 3
- Step-grate. JOSEF MARTIN Ger. 515,135 and 515,136, Dec 15, 1928
- Device for separating dust from gases from rotary kilns, etc. EISEN- UND STAHLWERK HORSCH & G. Ger 515,189, June 6, 1929
- Tilting plant for rotary drum ovens, etc. GEWERKSCHAFT EISENHÜTTE WESTFALIA Ger. 515,011, Jan 10, 1929
- Stamping device for use with rotary tube ovens. GEWERKSCHAFT EISENHÜTTE WESTFALIA Ger 515,392, Nov 21, 1929
- Fan for increasing the draft in annular kilns and for cooling the discharge chamber. GUSTAV SCHÜRDNER Ger 514,350, May 4, 1930
- Recording analyzer for industrial chemical control, more particularly of the purification of boiler-feed water. M. REIGNIER. Belg 370,021, June 30, 1930 A wheel, operated by a drum which is released at regular intervals, carries 4 hinged beakers. In position 1 the beaker is filled with the sample to be analyzed, in position 2 the titration is carried out and recorded, in position 3 the beaker is emptied, and at 4 it is refilled again. On the wheel are mounted spurs which open the cocks for the addition of the 3 reagents. A photoelec. cell opens the cock of the buret and closes it again when the phenolphthalein is decolorized
- Oil-feeding and -regulating device for oil burners. WM R. RAY (to Ray Burner Co.) U S 1,791,012, Feb 3 An automatic device is provided for varying the pressure on the oil in accordance with the temp. of the surrounding atm
- Gas burner. ALFRED J. REEDY (to A B Stove Co.) U. S. 1,791,247, Feb 3
- Gas burner. HARRY E. KERR (to Barber Gas Burner Co.) U. S. 1,791,360, Feb 3
- Bunsen burners. LOUIS MARTIN Fr. 693,535, April 8, 1930 Constructional details.
- Heat-exchange device suitable for heating or cooling liquids. HERBERT V. DALOLISSN (to Research & Development Corp.) U. S. 1,791,483, Feb 3 Structural features
- Heat-exchange apparatus suitable for use as a condenser, etc. STANLEY BROWN (to Grisco Russell Co.) U S 1,790,897, Feb 3 Structural features.
- Die-casting machine. A C SPARK PLUG CO Ger 515,140, April 19, 1929 Corresponds to Brit 332,277
- Centrifugal casting machine for making hollow bodies in horizontal molds. AURELIO POSSENTI and CARLO SCORZA Ger 515,139, Aug 14, 1929
- Suction box for felted-sheet-making machines. OSTILIO SEVERINI (to Società Invenzioni Brevetti Anonima-Torino) U S 1,791,065, Feb 3 Structural features
- Apparatus for continuous preparation of emulsions. RUDOLF AUERBACH U. S. 1,790,667, Feb 3 Structural features.
- Apparatus for sterilizing rags, etc., with hot air and steam. HEINRICH HELT Ger 515,405, Nov 2, 1927
- Acetylene generator. DAYTON H. DANLEY U S 1,791,104, Feb 3 Structural features
- Acetylene generator, with counter-pressure chamber 1 G FARNEMIND A-G Ger 514,437, Oct 8, 1929
- Rubberized conveyor belts R S CARTER (to Goodyear Tire & Rubber Co) Brit 335 475, April 16, 1929 The conveying surface is free from rubber so that it has a low coeff of friction while the driving surface is rubberized, and the belt may be formed of layers of fabric, some rubberized and some formed without rubber
- Photographic temperature-recording apparatus (suitable for use in ships' holds). OCEAN STEAM SHIP CO, LTD., and R H THORNTON Brit 335 644, July 6, 1929 Numerous structural details are described
- Thermostatic device. E L FONSECA (to Walcolator Co) Brit 335,819, Nov 27, 1928 Structural and mech details
- Thermostatic device suitable for control of electric circuits HAROLD D. EATON (to Bastian-Morley Co.) U S 1,791,477 Feb 3 Structural features
- Thermostatic electric switch. WM A RANKIN (to Robeson-Rochester Corp.). U S 1,791,225, Feb 3 Structural features.

2—GENERAL AND PHYSICAL CHEMISTRY

FREDERICK L. BROWN

- The life work of M. W. Beijerinck. JAN SMITH. *Chem. Weekblad* 28, 64-7 (1931).—A memorial address. G. G.
- The scientific career of Henry Louis Le Chatelier. RALPH OESPER. *J. Chem. Education* 8, 442-61(1931). F. H.
- Charles Edward Coates. W. L. OWEN. *Int. Eng. Chem.* 23, 339-40 (1931).—Biography with portrait. G. G.
- Reminiscences of Gibbs by a student and colleague. FOWEN B. WILSON. *Sci. Monthly* 1931, 211-27. E. H.
- Wilhelm Hofmeister. A. H. LARSON. *Plant Physiology* 5, 613-6(1930).—A biographical sketch accompanied by one plate. WALYFA THOMAS.
- Paul Julius. A. LÖTTINGHAUS. *Z. anorg. Chem.* 44, 109-12(1931).—Obituary with portrait. E. H.
- Håkan Sandqvist. ANON. *Svensk Farm. Tid.* 34, 633-4. Obituary with portrait. A. R. R.
- C. Willgerodt. O. HINSBERG. *Chem. Ztg.* 55, 85(1931).—Obituary. E. H.
- Nobel prizes in chemistry and physics for 1930. H. FISCHER and C. V. RAMAN. ANON. *Svensk Kem. Tid.* 42, 293-300(1930).—The accomplishments of these two scientists and their portraits are given. A. R. R.
- Unrecognized and unknown facts about the origin of the combustion theory of Lavoisier. MAX SFFIRK. *Chem. Weekblad* 28, 79-82(1931), cf. *C. A.* 20, 1922(German).—The influence of Guyton de Morveau and Mitouart on Lavoisier is described. F. SCHOTT
- Changing attitudes in the history of chemistry. J. N. SWAN. *J. Chem. Education* 8, 510-4(1931). E. H.
- An analysis of the college entrance examination board examinations in elementary chemistry for 1921 to 1923. JACOB W. BOORSTEIN. *Science Education* 15, No. 2, 91-100(1931). E. H.
- Questionnaire study of chemistry finances in South Dakota. RALPH DUNBAR AND J. ROBLEY LANE. *J. Chem. Education* 8, 565-7(1931). E. H.
- An administrative study on the teaching of chemistry in the secondary schools of South Carolina. G. G. NAIDAN. *J. Chem. Education* 8, 559-61(1931). E. H.
- An experiment on the use of journal articles in teaching elementary college chemistry. HOYT C. GRAHAM. *J. Chem. Education* 8, 531-3(1931). E. H.
- Dispensing laboratory solutions. CHARLES F. FICKELS. *J. Chem. Education* 8, 513-5(1931). E. H.
- Fertilizer from the air. A project for chemistry students at the secondary-schools level. HOWARD R. WILLIAMS. *J. Chem. Education* 8, 462-70(1931). E. H.
- An experiment on refractive index for the course in physical chemistry. CREGG S. HOYT. *J. Chem. Education* 8, 537-8(1931). E. H.
- Ashes and the remains of fuel found by the excavation of ruins of a seventeenth-century glass factory near the Keizergracht at Amsterdam. D. J. W. KREULEN. *Chem. Weekblad* 28, 97-102(1931).—The fuel used in this glass factory was a high-volatile coal of English origin. G. G.
- Geochemical distribution laws and cosmic frequencies of the elements. V. M. GOLDSCHMIDT. *Naturwissenschaften* 18, 999-1013(1930); cf. *C. A.* 24, 1259.—An elaborate review of the distribution of elements in the different spheres of the earth and in stellar bodies. It is concluded that the manner of distribution of the elements is detd. by properties of their electron shell; the frequency of occurrence of each type of atom, however, is detd. by properties of the nucleus. This frequency of occurrence is apparently the same in all parts of the cosmos. B. J. C. VAN DER HOEVEN
- Changeable mass and the second law. S. MONODONIC. *Archiv. Hem. Faem.* 4, 169-74(174-5 German)(1930).—Mathematical. J. KÜPERA
- The calculation of chemical, electrical and gas-mechanical transformations of matter, energy and form by space energetics. RICHARD VON DALLWITZ-WEGNER. *Z. Elektrochem.* 37, 25-33(1931).—On the basis of space energetics the equations derived for cohesion pressure, the processes in the Daniell cell, the mechanism of the galvanic cell, contact and thermal electricity gas laws, efficiency and entropy are discussed. H. STOERTZ
- Immersion liquids for determining refractivity of solid substances by the embedding method. A. MAYRHOFER. *Mikrochemie* 3, 52-71(1931).—Purpose. To establish a

standard series of immersion mixts. of a few liquids which could replace the numerous embedding fluids that have been proposed in the literature. The mixts.—water glycerol, cincole-paraffin oil paraffin oil- α -bromonaphthalene, α -bromonaphthalene- C_{11}H_7 —are considered standard mixts. of the first line. The following mixts. have high n and by diln with mixts. of lower n often serve for direct refractometric testing paraffin oil $\text{C}_{11}\text{H}_{18}$, C_{11}H_7 - α -bromonaphthalene, C_{11}H_7 , C_{11}H_5 . The percentage compns of the mixts. studied were detd. by careful measurement of the ingredients in microburets. The regularities with respect to the refraction corresponding to the mixt rule, the effect of keeping the mixts. for some time and the effects of temp were studied. W. T. HALL

Electrometric measurement of very low pressures by means of ionization currents the space charge manometer. WERNER MOLLATH *Z tech Physik* 11, 522-9(1930)—By the use of an electrometer method the sensitivity of pressure measurement by ionization can be raised from 10^{-1} mm Hg to as low as 10^{-10} mm Hg. A measuring tube with 4 electrodes is used in a circuit with electrometer and 2 tnodes, calibration method and details are given. B. J. C. VAN DER HOFVEN

The development and present state of our knowledge of the natural system of elements (eulogary of Lothar Meyer's birthday). F. PANETH *Naturwissenschaften* 18, 964-76(1930)—A review. B. J. C. VAN DER HOFVEN

The ozone of the upper atmosphere. CH. FANAY *Scientia* 49, 11-22(1931)—This is a summary of the evidence relating to the presence and manner of distribution of ozone in the atm. The evidence is derived principally from the absorption spectra of the sun and stars. If the ozone in the atm. were uniformly distributed its concn. would be 60 mg per 100 kg of air. Actually, at the earth's surface, air contains about 3 mg of ozone per 100 kg. Most of the ozone is distributed in the upper regions of the atm at altitudes which are variable, and which are estd. to be from 20 to 50 kilometers. If the atm. ozone were segregated it would form a layer about 3 mm. thick at atm. pressure. The thickness of this layer, however, varies between extreme limits of 2 and 4 mm due to seasonal and irregular causes. The origin of atm. ozone is uncertain. Speculations are advanced regarding the relation of atm. ozone to certain phenomena of the upper atm. It is estd. that atm. ozone is responsible for the absorption of about 4% of the sun's energy that is incident on the upper atm. R. H. L.

Unit of atomic weight. F. W. ASTON *Nature* 126, 953(1930)—None of the suggestions so far advanced for a new unit of at. wt. is free from objections. While extreme accuracy is not vital for chemists, physicists now need an accuracy of 1 in 10 000. GREGG M. EVANS

Magnetic rotary power of hydrocarbons in the gaseous state. R. DE MALLEMANN AND P. GABIANO *Compt rend* 191, 1319-20(1930)—By using $\lambda = 578 \text{ m}\mu$, the Verdet const., sp. and mol. rotations for methane, ethane, propane, butane, isopentane and hexane were detd. for the gaseous state and, for the 3 latter, in the liquid state. The dispersion ratios for $\lambda = 546 \text{ m}\mu$ and $\lambda = 436 \text{ m}\mu$ are 1.125 and 1.84 for all the hydrocarbons. F. W. LAIRD

The magnetic rotatory polarization of fused organic substances. CONSTANTIN SALCEANU *Compt rend*, 191, 1316-8(1930), cf *C. A.* 24, 3682—The variation with temp. of the magnetic rotatory power and magnetic double refraction of fused naphthalene and β -methylnaphthalene was detd. for 3 Hg lines. There is a slight convexity toward the temp. axis of the rotation temp. curve which is the opposite of that obtained for the magnetic double refraction temp. variation. The experimentally detd. values for the magnetic rotary dispersion agree fairly well with those calcd. from Mallmann's theory. F. W. LAIRD

Magnetic double refraction of phenol, naphthalene and phenanthrene in the molten state. CONSTANTIN SALCEANU *Compt rend*, 191, 480-8(1930)—By means of the exptl. arrangement of Cotton and Dupouy (cf *C. A.* 24, 3682) the magnetic double refraction of phenol at 49° was found to be 1.10° , of naphthalene at 106° , 3.45° , and of phenanthrene at 104° , 7.83° . The double refraction is given at several other temps. for each compd., as are the ratios of the double refraction to that for nitrobenzene. The Cotton Mouton consts. also are given. E. G. VANDEN BOSCH

Molecular field and atomic order in ferromagnetic crystals and in hydrogenated iron. L. W. MCKEEHAN *Nature* 126, 952-3(1930)—"The 2 possible ways in which magnetization may change are (1) by reversals, without sensible magnetoresistance changes or magnetostriiction, and with small but definite hysteresis losses probably closely conditioned by elec. cond., and (2) by rotations through less than 180° , with magnetoresistance and magnetostriiction, with hysteresis losses of important amount

largely controlled by mechanical strains inherent or induced by applied stresses, and possibly subject to catalytic acceleration. GRIGG M. L. A. 1929A.

Generalization of the virial of Clausius. S. RAY. *Bull. intern. acad. polonaise* 1929A, 233-9—Math. The virial of Clausius is applied to the case of a vol. of gas enclosed within impervious walls disposed symmetrically with respect to the co-ordinate axes. B. C. A.

Entropy, elastic strain and the second law of thermodynamics; the principle of least work and of maximum probability. W. S. KIMBALL. *J. Phys. Chem.* 35, 611-23 (1931) cf. *C. A.* 23, 5359. A new mech. aspect of entropy with reference to a simple, fully excited gas is obtained from the geometrical expression for wt. and by taking strains to include unit extensions in velocity and momentum space as well as ordinary space. The second law of thermodynamics and the automatic increases of entropy that it represents are attributed to increased strain under the action of corresponding stresses, rather than to the unsatisfactory ergodic hypothesis as in statistics.

The nature of the effect of the addition of salt upon the surface tension of sodium taurocholate solutions. MASAE SHONO. *J. Biochem. (Japan)* 12, 317-39 (1930)—Various salts added to a soln. of Na taurocholate lower its surface tension to different degrees. At higher concns. the effect is conditioned upon the hydration of the salts while at much lower concns. the effect depends upon the nature of the cations, those forming less dissociable salts being more effective in depressing the surface tension of the Na taurocholate solns. The surface tension of taurocholic acid or of its salts is due to the accumulation of their mols. at the surface of soln. The min. surface tension of the Na, Ca and Cu salts differ from each other, indicating that the surface activities of the films of these various taurocholates differ. S. MORIGUCHI.

Binary mixtures. L. FIATTE. *Z. physik. Chem. A*, 152, 39-46 (1931)—The b.p. compn. and the viscosity-compn. diagrams at various temps. were detd. for mixts. of *m*-cresol and PhOMe with $\text{C}_6\text{H}_5\text{OH}$ and $\text{C}_6\text{H}_5\text{I}$. At 0° the viscosity curves are nearly straight lines but at higher temps. the deviation is marked. The results are discussed with respect to mol-compd. formation. V. F. HARRINGTON.

Maximum velocity of evaporation of liquids evaporated on heated metallic surfaces. ZORYA BLAZIKOWSKA. *Rozniki Chem.* 10, 691-713 (713-4 French) (1930)—The phenomenon of max. velocity of evapn. of liquids was investigated by evapn. 10 liquids (Et_2O , Me_2CO , CHCl_3 , MeOH , EtOH , CCl_4 , C_6H_6 , H_2O , $\text{C}_6\text{H}_5\text{Cl}$, and $\text{C}_6\text{H}_5\text{I}$) over Al, Fe and brass surfaces heated several tenths of a degree above the boiling temp. On heating the metallic surfaces to temps. still higher a rapid diminution of the velocity of evapn. can be observed, this is due to the Leidenfrost phenomenon, which makes the evapn. very difficult. A small elec.-furnace was constructed which served to study and det. with precision the so-called "temp. corresponding to the max. velocity of evapn.," that is, the temp. of the metallic surface at which the velocity of evapn. attains a max. A series of curves is drawn showing the relation between the time of evapn. of 1 drop of the liquid (approx. 0.01 g.) and the temp. of the metallic surface on which the evapn. takes place. The curves are marked by a min. time of evapn. of the liquid (corresponding to the max. velocity of evapn.) in a certain temp. interval lying above the b.p. of the liquid. On the basis of these exptl. results a technical method is elaborated for a very rapid evapn. of liquids by heating the metallic surfaces to the "temp. corresponding to the max. velocity of evapn." 600 l. of C_6H_6 or $\text{C}_6\text{H}_5\text{Cl}$ could be evapd. in one hr. per 1 sq. m. Also in *Bull. intern. acad. polonaise* 1930A, 188-90.

The rate of shear in capillary tubes. WINSLOW H. HERSHEL. *J. Rheol.* 1, 505-7 (1930)—A brief math. note suggested by Williamson's paper (cf. *C. A.* 24, 5193).

Rheology index. EUGENE C. BINGHAM. *J. Rheol.* 2, 10-107 (1931).—A list of 5000 references to rheological literature arranged alphabetically according to authors. The subjects of papers are included and in many cases the location of abstracts of papers are cited. EUGENE C. BINGHAM.

Some fundamental definitions of rheology. EUGENE C. BINGHAM. *J. Rheol.* 1, 507-17 (1930)—This is a revised form of definitions drawn up for the comm. on plasticity, consistency, etc., of the Am. Soc. for Testing Materials (cf. *Proc. Am. Soc. Testing Materials* 29, 11 (1929); *C. A.* 23, 4119, 4861).

Plasticity and criteria of molecular aggregation. S. E. SHEPPARD. *J. Rheol.* 1, 471-84 (1930); cf. *C. A.* 24, 1783—S. discusses the work of Staudinger and his co-workers, Whitby and his co-workers, Kraemer and Williamson, Filkentscher and Mark and attributes the solvation of colloids to monomol. adsorbed layers formed by dipole

checked those obtained by Cox (*C. A.* 22, 4296) 0.771 1 0.704. These values are near those of thiophene, 0.757 1 0.757, while the vols. of the unit cells are 471 and 498×10^{-24} cc., resp., which would explain the formation of solid solns. of these 2, although there is no true isomorphism. A. W. COVITTA

Lattice constants of mixed crystals of copper-nickel. W. G. HURGER AND J. C. M. BASART. *Z. Krist.* 75, 155 7 (1930).—Precision measurements on a series of mixed crystals of Cu-Ni show a slight deviation from the values as predicted by Vegard's additive law. The max. deviation was 0.008 Å. L. S. RAMSDELL

Anomalous crystallized mixed systems of the type iron ammonium chloride. A. NEUBAUER. *Chem. Erde* 5, 551 621 (1930). NH_4Cl may take up varying amts. of $\text{FeCl}_2 \cdot \text{H}_2\text{O}$ (up to 7.5%) giving distorted cu. crystals showing optical anomalies. Crystals of a similar nature are formed with $\text{CoCl}_2 \cdot 2\text{H}_2\text{O}$ (up to 19%) and $\text{NiCl}_2 \cdot 2\text{H}_2\text{O}$ (up to 14%). Conclusion: These are not double salts nor true mixed crystals, but rather represent the inclusion of foreign matter along certain faces in the NH_4Cl crystals. The systems $\text{FeCl}_2 \cdot \text{H}_2\text{O}$, $\text{CoCl}_2 \cdot \text{H}_2\text{O}$ and $\text{NiCl}_2 \cdot \text{H}_2\text{O}$ were also studied in this connection, and data are given for some new crystal hydrates. B. C. A.

Artificially colored crystals as examples of the so-called anomalous mixed crystals. K. SPANGENBERG AND A. NEUBAUER. *Chem. Erde* 5, 417 528 (1930).—A detailed review is given of artificially colored mixed crystals (e. g., FeCl_3 in NH_4Cl , methylene blue in $\text{Pb}(\text{NO}_3)_2$, etc.) and of the influence of such substances in soln. on the habit of the crystals. It is considered that these are not true mixed crystals but represent an inclusion of foreign material adapted to the structure of the host. In this connection the composite crystals formed by the regular intergrowths of various minerals and salts are considered in detail from a structural viewpoint. B. C. A.

Distribution of foreign substances in crystals. J. NOTENHAFT AND H. STEINMETZ. *Chem. Erde* 5, 225-32 (1930).—The distribution of color and inclusions in various minerals and in crystals of $\text{Pb}(\text{NO}_3)_2$ grown from solns. contg. various coloring matters is described. In fluorite the bands of color parallel to the cube faces are built up of minute fibers perpendicular to the cube faces, and there are diagonal areas in the crystal free from inclusions. The observations are discussed in connection with the growth of crystals. B. C. A.

Anomalous crystallized mixed systems with organic components or of an organic component in an inorganic host. A. NEUBAUER. *Chem. Erde* 5, 529-53 (1930).—Various inorg. salts when crystal. from soln. may take up org. coloring matters; e. g., $\text{Ba}(\text{NO}_3)_2$, $\text{Pb}(\text{NO}_3)_2$, and $\text{Sr}(\text{NO}_3)_2$ readily take up methylene blue, NaCl and KCl take up murexide, and K_2SO_4 takes up ponceau red and Bismarck brown. Similar anomalous crystals are also formed by certain pairs of org. compds., e. g., C_{10}H_8 and dihydronaphthalene. Crystals of this kind are considered in detail especially from a structural point of view. B. C. A.

The points on the isotherms of electrical conductivity of metallic mixed crystals. G. GRUBE AND J. HILLE. *Z. anorg. allgem. Chem.* 194, 170-89 (1930).—The relations between the constitution of binary alloys and their elec. cond. are so far explained that this property may be used to obtain constitution diagrams in cases where thermal analyses do not yield clear results. If a non-dissoc. intermetallic compd. seps. from a continuous series of mixed crystals of 2 metallic components which in turn forms mixed crystals with the components, no displacement of the max. of the cond. isotherm occurs with temp. if both components lower the transformation temp. of the compd. If the transformation temp. is increased by both components, 2 rounded max. occur on the isotherm at temps. adjacent to the transformation temp. of the pure compd., which lie at lower and higher temps. than the compd. If the transformation temp. is increased by one component and lowered by the second, there may occur a flattening and displacement of the max. to other temps. at adjacent temps. The heterogeneous decompn. regions adjoining the α - and β mixed crystals have the same cond. at the same temp. For the establishment of the phase boundaries, the temp. resistance curve rather than the cond. isotherm is used. A displacement of the max. of the cond. isotherm with the temp. may also occur if both components lower the transformation temp. of the compd., only however if the compd. is partly dissoc. into its components. ALLEN S. SMITH

Conditions of growth of crystals of difficultly soluble substances. I. D. KURBATOV. *Compt. rend. acad. sci. U. R. S. S.* 1930A, 429-35 (1930).—Conditions of accelerating or decelerating the process of crystal. of difficultly sol. substances were studied, and a distinction was drawn between the methods employed in the labs. and those natural conditions that give rise to the growth of large crystals. A new lab. method for obtaining large crystals, based on exptl. evidence and theoretical considerations, is pro-

posed. Large crystals can be obtained best by using solns. of non-equiv. ionic concns. of the substance to be crystd. To feed crystals formed in satd. solns. and to prevent formation of new centers of crystn. non-equiv. solns. should be used. They present 2 factors favoring growth of large crystals: (1) increase of the soly. product, and (2) decrease of the no. of mols. that can crystallize out of the soln. B. N. DANTLOFF

Magnitudes of state of oxygen at low temperatures. FRITZ SCHMIDT, *Z. Ver dent Ing* 74, 1710(1930).—The thermodynamic properties of O at temps. between $+200^{\circ}$ and -200° and at pressures from 1 to 1'0 atm. were detd. by interpolating and extrapolating isothermic curves. The following values for the most important properties were derived:

	Temp.	Pressure	Sp. Gravity γ in kg./dm. ³
N l p	-218.4	1 mm Hg	1.305
l p at atm. pressure	-182.8	760 Hg	1.141
Crit. point	-118.8	51.35 atm.	0.450

The sp. heats were derived from the pe diagrams: sp. heat of liquid $c_p = 0.397$ kg.-cal./kg. $^{\circ}\text{C}$ at -200° and the heat of vaporization $r = 51$ kg.-cal./kg. at -182.8° and 760 mm Hg. M. McMANIS

The compressibility of nitrogen and hydrogen at pressures up to 5000 atmospheres. JAMES BARRETT AND R. DUFFIN, *Compt. rend.* 191, 1295-7(1930).—The vols. of H_2 and N_2 that occupy 1 cc. at 16° and 1000 kg. per sq. cm. are given at pressures ranging from 1000 to 5000 kg. per sq. cm. Similarly the vols. of H_2 and N_2 that mutually occupy 100 cc. at 0° and 76 cm. are given for these same high pressures. The values for H_2 are in accord with those of Amagat. App. for subjecting gases to 10,000 kg. per sq. cm. and to 25,000 atm. is under construction. F. W. LAIRD

The compressibility isotherms of methane at pressures to 1000 atmospheres and at temperatures from -70° to 200° . HAMLINE M. KVALNES AND V. L. GIBBY, *J. Am. Chem. Soc.* 53, 3449(1931).—Published data were checked and addnl. values obtained to cover 8 temps. and 25 pressures at each temp. B. A. SOTER

A method for the determination of critical temperatures and the critical temperature of hydrogen fluoride. P. A. BOND AND DELLA A. WILLIAMS, *J. Am. Chem. Soc.* 53, 34-7(1931).—The crit. temp. of H_2F is 230.2° , as detd. by a new method in which the pointer of a balance is set in motion upon the appearance of the liquid phase. The method, which may be used for substances attacking glass, has been checked (a) directly on SO_2 , (b) by substituting in van Laar's equation $T_c/T_b = K$, values from H_2F and obtaining for K a value (1.72) very close to that for HCl , HBr and HI . DON BRODER

The activation energy of adsorption processes. HUGH S. TAYLOR, *J. Am. Chem. Soc.* 53, 579-97(1931), cf. C. A. 25, 621.—Adsorption is not necessarily rapid. The assumption that adsorption processes possess their own characteristic activation energies leads to a general theory of adsorption that applies equally well to the so-called primary, or irreversible, and to the secondary, or reversible, adsorption. Those adsorptions are secondary or reversible that have small heats of adsorption and small activation energies. In general, it is the adsorptions with the higher energies of activation that are important in catalytic change. F. L. BROWNE

Velocity of adsorption processes and the problem of promoter action. HUGH S. TAYLOR AND ARTHUR T. WILLIAMSON, *J. Am. Chem. Soc.* 53, 813-4(1931), cf. preceding abstract.—On an active MnO catalyst H_2 is adsorbed immeasurably slowly at 0° and 100° , at 180° the initial rate of adsorption is 8 mm. per hr. and at 305° 78 mm. per hr., corresponding to an activation energy of approx. 10,000 cal. A $\text{MnO}-\text{Cr}_2\text{O}_3$ catalyst shows an extremely low rate of adsorption at 0° but at 184° it has risen to approx. 100 mm. per min. or an increase of about 700-fold by reason of the promoter action of the Cr_2O_3 . In considering the mechanism of promoter action attention must be paid not only to extension of surface but also to the velocities with which the activating adsorptions occur. C. J. WEST

Surface films. N. K. ADAM, *Nature* 126, 955-6(1930).—A comment on a review by Langmuir of A's book "The Physics and Chemistry of Surfaces" (cf. C. A. 24, 2354). A disclaimer as much disagreement with L. as is inferred by the latter. The liquid expanded state is intermediate between the condensed and the expanded states, and it is to this state only that the statements as to competition between the chains and heads of mols. apply. A. however, reaffirms his belief in the possibility of definite angles of tilt. GREGG M. EVANS

Dialysis, ultrafiltration, osmosis and their applications. A BOUTARIC. *Chimie & Industrie* 24, 1295-303(1930)—A review A PAPINFAU-COUTURE

Finely divided solids as emulsifiers. J L VAN DER MINNE. *Chem Weekblad* 27, 651-2(1930)—The conclusions of Carrière (C. A. 25, 856) are compared with those of M. (Dissertation Utrecht, 1928) cf. C. 1 22, 2312 B J C VAN DER HORSTEN

A method for measuring average particle size of emulsions. WUFELER P DAVEY. *J Phys Chem* 35, 115 7(1931).—When a drop of an oil in water emulsion falls into clean neutral water with force enough to break the surface of the water, the drop retains its identity for some time. A drop of emulsion deposited gently on the water spreads very rapidly. When spreading is complete a layer one colloidal particle deep seems to exist on the surface of the water. If it is assumed that the colloidal particles remain spherical, the surface covered by a known quantity of emulsion gives the information necessary for calculating the diameter of the particles. The emulsion must be dilute enough to avoid aggregation of particles and covered enough to furnish sufficient particles to cover the surface. Samples of an emulsion at different dilutions indicate the same particle size. The procedure is the same as that employed by Langmuir with oil films.

J. I. BROWN

Influence of electrolytes on the dispersion of clays. F ALLAM. *Chem Erde* 5, 276-318(1930).—The results are given of a no. of expts. on the permeability and flocculation of a clay from Silesia with various salt solutions. B C A

Liquid ammonia as a lyophilic dispersion medium. II. Ammonio-gels of cellulose acetate. ROBERT TAFT and JESSE E. STARECK. *J Phys Chem* 35, 678-87(1931).—Sols contg. as much as 60 g cellulose acetate in 100 cc. of liquid NH₃ were prepd. A mixt. of acetates is more readily dispersed than the cellulose triacetate. The sols of higher concn. were prepd. by evapn. of part of the solvent. There seems to be no limit to the possible concn. of sol. Dil. sols are clear, colorless and mobile, concd. sols are viscous, translucent and light brown. On warming, the viscosity decreased at first but after several hours the sols changed irreversibly to white opaque gels. The time required for gelation is a function of concn., temp., H₂O content and age. Gels were formed by cooling sols to about -60°. These gels were reversible on warming to a temp. somewhat above that of their formation. Heat irreversible gels seemed to have changed chemically and to have adsorbed water. F I. BROWN

Determination of the hydrogen-ion concentration in gold sols. T R. BOLAN AND J CAOWE. *J Phys Chem* 35, 102-10(1931).—By using a specially designed H⁺ electrode requiring only a little soln., it was found possible to det. the pH values but the sol must be either sufficiently acid or suitably buffered in order to obtain steady and reproducible results. Probably there was some foreign material present during the expts. of Adolf and Pauli, and of Tartar and Lorah, who were unable to obtain satisfactory values. W. T. H.

Solution of colloidal particles by dilution of the sol. ADOLPH J RABINOVICH AND V. A KARGIN. *Z. physik Chem., Abt. A*, 152, 24-35(1931).—The previously discovered property of acidoid sols of dissolving on diln. is confirmed with WO₃ and V₂O₅, and the results were checked by electrometric and spectrographic measurements. Because of the formation of H⁺ ion by hydration of the inner layers of the particles, these sols behave similarly on diln. of their buffered solns.; their pH remains const. until the particles are in soln. The Tyndal effect then disappears. The not purely acidoid sols, As₂S₃ and Fe₂O₃, do not behave thus. V. F. HARRINGTON

Polystyrene oxy-combinations that form during the synthesis of electronegative hydrosols. V. Oxyacids. A. V DUMANSKI and A. G. YAKOVLEV. *Bull. soc. chim.* [4], 47, 1211-6(1930), cf. C. A. 24, 3939, 4977.—Triangular diagrams for the sols of the systems comprising FeCl₃ 0.1 N [A], NaOH and NH₄OH 0.1 N [B], and the Na salts of acetic glycolic, succinic, malic, tartaric and citric acids 0.1 N [C], are worked out. The oxyacids give peptization diagrams with regions of deposition, of cloudy electroneg. sols, of electroneg. sols without pptn., and of electropos. sols with and without pptn. Equil. is reached very slowly in these systems. The peptizing action is of the order: tartaric acid > glycolic acid > malic acid > citric acid. Increasing the OH groups increases the peptization properties as shown by tartaric, malic and succinic acids. Substituting NH₄OH for NaOH increases the peptizing action. Thirteen diagrams are shown. Also in *J. Russ. Phys.-Chem. Soc.* 62, 1665-70(1930).

H. W. WALKER

The mechanism of formation of colloidal silver. HELEN Q WOODWARD. *J. Phys. Chem.* 35, 425-31(1931).—When Ag sols are formed by the Bredig method, Ag reacts with the anion present to form salts, complexes or AgOH. Expts. are reported for arc-

ing in H_2O and in solns. of 14 electrolytes. When the Ag salts form, the cations of the stabilizing salt form bases. In water $AgOH$ forms. F. E. BAOWE

Chromium hydroxide. JELIUSZ LISIECKI. *Poczniki Chem* 10, 736-42 (1942 German) (1937).—A Cr_2O_3 sol with as small amt. of peptizing agent as possible was prepd. in a pure and concd. form without dialysis. The excess of peptizing agent can be removed by satn. with a corresponding amt. of gel, in the form of dispersoid instead of by dialysis. The time of prepn. is shortened from several weeks to several hrs. and the exptl. conditions are quant. By this method various metal oxide hydrosols of any concn. can be prepd. with any ratio of the peptizing gel to peptizing agent. A metal oxide hydrosol differs from the dispersoid although it does belong to the disperse systems. The dispersoids show different cataphoretic properties as compared with hydrosols and are not pptd. by electrolytes. As Bredig's expts. have clearly proved, stable hydrosols cannot be prepd. without peptizing substances. Besides mechanical or elec. dispersion, chem. dispersion or peptization is necessary to prep. a permanently stable hydrosol. J. KUČERA

Hydrosol of silicic acid. I. Hydrosol made from ethyl silicate by hydrolysis. KENYU INADA. *Sci. Papers Inst. Phys. Chem. Research* (Tokyo) 15, 1-15 (1939).—Unless the ethyl silicate, prepd. from $SiCl_4$ and $EtOH$ and purified by vacuum distn., contains some HCl , the H_2SiO_3 hydrosol resulting from hydrolysis is unstable. Hydrolysis of pure ethyl silicate in the presence of such peptizing agents as $NaOH$, KOH , HCl and H_2SO_4 (but not H_2CO_3 or org. acids) gives stable sols. The 2 types of sols studied were prepd. by hydrolysis of ethyl silicate in the presence of (1) HCl and (2) $NaOH$. Each sol was purified by electro dialysis through a parchment membrane. The highest concn. of SiO_2 that can be stabilized by HCl is 7-8%. The HCl cannot be removed completely without causing coagulation, e. g., a sol contg. 2.66% SiO_2 becomes unstable when the HCl concn. is reduced below 0.001 N. Electro dialysis of $NaOH$ -stabilized sols contg. more than 14% SiO_2 gives unstable sols, with less SiO_2 present strictly neutral sols of high stability can be produced. Because of the relatively high elec. cond. and low viscosity of the neutral sols, the particles must be smaller in the neutral than in the acidic type of H_2SiO_3 hydrosol. This is confirmed by the fact that the amt. of SiO_2 passing through the parchment during electro dialysis is negligible in the acidic but very large (30-50%) in the neutral sols. The stability of the acidic sols varies with the HCl and SiO_2 concns. Addn. of small amts. of HCl to the neutral sols tends to increase the stability. II. Electrical properties of the hydrosol of silicic acid. *Ibid.* 16-31.—The 4 sols used differed in method of prepn.: A was the acidic and B the neutral sol described in the preceding abstr.; C was made by hydrolysis of $SiCl_4$; D was made by dissolving in $NaOH$ soln. some of the H_2SiO_3 sol from hydrolysis of $SiCl_4$ and then decomposing the sodium silicate with HCl . Each sol was purified by electro dialysis. Sol C resembled A in requiring some HCl to stabilize it, although D could be obtained in neutral form, not all of the $NaCl$ could be removed. Electro phoresis measurements were made in a U tube of the Burton type (diam. of tube 1.1 cm. and distance between electrodes 23 cm.). The movement of the boundary which was practically invisible to the naked eye, was detd. by photography in ultra-violet light, with an ordinary plate. Under a potential gradient of 1 volt/cm. at 25° the migration velocities $\times 10^4$ of the boundary (toward the anode) in cm./sec. were 2, 23, 1 and 11 for sols A, B, C and D, resp. each contg. approx. the same amt. SiO_2 (0.3-0.4%). The corresponding conductivities (units not specified) were 1.97×10^{-4} , 0.675×10^{-4} , 1.15×10^{-4} and 0.688×10^{-4} . Variation of the HCl concn. (c) revealed a max. migration velocity (x) at low c in sols A, B and C. With increasing c in sols B and D, x decreased to zero and reversed sign. In sols A and C the max. x was followed by an abrupt decrease and then a gradual decrease through zero. The effect of KCl , $BaCl_2$, $CeCl_3$ and $ThCl_4$ on neutral sol B was studied. The effect of KCl and $ThCl_4$ upon x was similar to that of HCl . With $BaCl_2$ coagulation took place before an isoelec. point was reached. At concns. of 0.0003-0.0004 N $CeCl_3$ the H_2SiO_3 moved toward the cathode but outside these limits the sol moved toward the anode. The viscosity of the sol shows a min. in the neighborhood of the isoelec. point. The adsorption of neg. ions Cl or silicate is as important a stability factor as the charge or hydration and its effect is most marked near the isoelec. point. OSCAR T. QUIMBY

The syneresis of silica gel. J. FERGLSON and M. P. APPLEBY. *Trans. Faraday Soc.* 26, Pt. II (42), 579-80.—The syneresis of SiO_2 gel was studied under several conditions. The vol. of syncretic liquid expressed exactly equals the decrease in vol. of the gel. The vol. of liquid when plotted against time gave an S-shaped autocatalytic curve. The velocity of syneresis and the total vol. of liquid expressed were greater the greater the original SiO_2 concn. of the gel at the time of setting. The

velocity of synthesis was approx. doubled for each 10° rise in temp. The initial velocity of synthesis was greater the more alkaline the gel, but the total vol. finally expressed was greater the more acid the gel. The hydration of the micelles was one mol. of water to one mol. of SiO_2 in neutral gels. The hydration increased with alk., being two mols. of water to one of SiO_2 at a pH 10. It is concluded that the progress of synthesis is aided by two factors: (1) the tendency of SiO_2 micelles to consolidate further by the action of unreacted residual valences, and (2) the tendency of the structure to resist deformation. I. H. REYERSON.

Coagulation of ferric oxide hydrosols. C. HANSEN-SØRENSEN. *J. Am. Chem. Soc.* 53, 812 (1931). Solenized with Drier (C. J. 24, 6572). C. J. WEST.

Pseudo- "irregular series" observed in colloidal gold solution. LUCIEN LEWIS. *Bull. Inst. Phys. Chem. Research* (Tokyo) 9, 641 (1930). (*Abstracts* 89 (in English) published with Sci. Papers Inst. Phys. Chem. Research (Tokyo) 14, Nov. 25 '30.) A colloidal Au soln. is prepd. by adding 10 cc. AuCl_3 soln. (0.04% $\text{AuCl}_3 \cdot \text{HCl} \cdot 3\text{H}_2\text{O}$) and reducing agent to 100 cc. dist. water stirred well at room temp. The soln. is heated and when it begins to boil KOH or HCl in various concns. is added to det. the influence of the acids on the stability of soln. As the reducing agent the following 5 compds. are selected: 3.5 $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$ (I) 5 cc. of soln. prepd. by adding 10 cc. formalin to 90 cc. 0.2 N KOH soln. (II), 3.5 $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$ (III) 6 $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$ (IV), and 8 $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$ (V). When (I) is used, the colloidal soln. is stable as long as the concn. of alkali added is less than 1.2 millimol./l. and more than 5.6 millimol./l. The coagulation value of KOH is 21 millimol./l. In (I) and (II) systems the second stable region exists, in general, in the sphere of 5-8 millimol./l. and 4.5-6.5 millimol./l. In (IV) system the soln. has comparatively a good stability when there is 6 millimol./l. of KOH in the ultimate concn. In (V) it shows a more tendency to become stable. The soln. of HCl shows no special phenomenon. In each colloidal soln. coagulation values by H_2SO_4 and NaCl and pH are resp. measured. Though the phenomenon resembles the so-called "irregular series," it differs from it in that (1) colloidal particles in the second stable region seem to have the same neg. charge as those in the first stable region and (2) the phenomenon is observed only when the soln. is cooled out in the course of the reaction of prepn. K. KONNA.

Distribution of hydrogen and hydroxyl ions in gelatin cubes. R. J. HANCOCK. *Compt. rend. soc. biol.* 102, 100-1 (1929), *Physiol. Abstracts* 15, 11.—Gelatin cubes containing phenol red were placed in solns. of weak NaOH plus phenol red. In 24 hrs. each presented a yellow center and a red peripheral zone, the thickness of which was the same in all blocks of different sizes. The pH (8 to 10) of the soln. directly changed in the course of a week. In the second week the pH changed and in the blocks the acidity increased from the surface toward the center. The values of pH in the cubes were comparable only when measured at corresponding distances from the surface. G. G.

Distribution of hydrogen and hydroxyl ions in gelatin columns. R. J. HANCOCK. *Compt. rend. soc. biol.* 102, 102-1 (1929), *Physiol. Abstracts* 15, 11-2.—The phenomenon described in the preceding abstract were studied in the case of gelatin placed in tubes. In the case of agar jelly the red color reaches the bottom of the tube in 2 days. Agar does not react chemically with penetrating electrolytes. In the gelatin tubes a colorless portion remains at the bottom 25 days later. G. G.

Role of structure of jelly in permeability of ions. R. J. HANCOCK. *Compt. rend. soc. biol.* 102, 524-6 (1929); *Physiol. Abstracts* 14, 563-7.—The selective permeability of gelatin for cations and anions varies with the pH of the liquid. The permeability depends on chain structure, and not on properties of a membrane of sepiol , as suggested by Girard. Protein presents the properties of an active ion. In the setting up of a Donnan equl. it is unimportant whether the particles of protein be grouped in layers or scattered in the gelatinous phase. The eventual role of structure is indirect; it contributes to the semi-solid consistency of the gelatin and to the fixation of the nondiffusible protein ions. G. G.

Abnormal osmosis at non-swelling membranes. I. K. SÖRENSEN. *Z. Elektrochem.* 36, 39, 40 (1930).—The magnitude of the p.d. between two solns. of electrolytes sep'd. by a cellulose membrane depends on the previous treatment of the cellulose, whereby the size of the pores is affected. Such membranes contain fine parts of wholly differing dimensions, thereby causing potentials to be set up which differ from plate to plate and from pore to pore and in consequence tend to establish closed current circuits within and around the membrane. B. C. A.

The freezing of gel. KYOGI KYOSHITA. *Bull Chem Soc Japan* 5, 261-6 (1930) — An app is described in which sols of gelatin were frozen. Cooling curves were obtained for sols varying in concn from 5 to 60% gelatin. Undercooling was pronounced and is attributed to inhibitive action of gelatin on the formation of ice crystals. Alternate ice and gel layers were found in sols of concn up to 15% gelatin. Up to this concn the freezing temp fell less than 3°. In the range of sols between 15 and 30% the freezing temp fell rapidly to -0.74°. It is concluded that there is almost no water in this region that can be sepd by freezing. With still greater concn of gelatin, water must be more firmly bound, since the f p is not depressed as rapidly. The f p-concn curve is of an inverted S type. L. H. REYERSON

The vapor pressure of gels. F. H. BUCHNER. *Compt. rend.* 191, 1323-4 (1930) — Two tubes of gelatin gel were placed with their lower extremities in water. In one case the upper portion of the gel was exposed to satd water vapor and in the other the end of the tube was closed. After one month the upper portions of both gels appeared firm and analysis showed that the upper extremity in each case contained more gelatin than the lower. Further, the final content of the upper ends was less than the original, indicating absorption of water in both cases through ascension. F. W. LAIRD

The vapor pressure of gels. PAUL BAAY. *Compt. rend.* 191, 1325-6 (1930); cf *C. A.* 24, 4302 — A controversy with M. Buchner concerning the vapor pressure of a gel and its swelling liquid. F. W. LAIRD

The denaturation of albumin. WILDER D. BANCROFT AND J. E. RUTLER, JR. *J. Phys. Chem.* 35, 141-61 (1931) — It has been assumed that proteins are irreversibly coagulated and that this process is really denaturation followed by, or accompanied by, coagulation. Denaturation was supposed to be a chem change. It is found that there really is no denaturation and Al, KCN, $(\text{NH}_4)_2\text{CO}_3$, NH_4CNS , NaHCO_3 , CH_3O and cane sugar pepture heat coagulated egg white sols. The peptized sols produce the same immuno-biological test as the original egg white, and have the same isoelec. points. 1 ther exts from egg white sols a substance that acts like crude lecithin. After this substance was extd from egg white sol, immersion in boiling water for 15-20 min failed to produce coagulation or cloudiness in a 10% sol. The coagulation and peptization of egg white are reversible colloidal changes. An extensive literature review is included. F. P. BAOWY

Scattering of light in protein solutions. I. Gelatin solutions and gels. K. KRISHNAMURTI. *Proc. Roy. Soc. (London)* A129, 490-509 (1930), cf *C. A.* 24, 1579 — Gelatin sols are considered polydisperse systems in which part of the gelatin is present in the molecularly dispersed condition, and the rest as polymol micelles. Cooling below 25° produces superatn, as a result of which fresh particles are formed. These become large as a result of the condensation of the molecularly dispersed gelatin on them. During the gel sol transformation the gel micelles are gradually dispersed in the original mol condition. The variation of the intensity of the scattered light with the concn of gelatin sols and gels was examd. The depolarization of the scattered light first decreases and then increases when gelatin sols are cooled to 10°. The significance of this in revealing the change in the size and shape of micelles is pointed out. The turbidity at the isoelec. point is caused by the aggregation of gelatin mols. F. U.

The colloidal structure of egg white as indicated by plasticity measurements. J. L. ST. JOHN AND F. L. GARF. *J. Rheol.* 1, 494-505 (1930), cf *C. A.* 24, 5090 — Egg white consists of a thin portion 32% and a thick portion 68%. The plastometer of St. John was used to study the apparent fluidity of both portions. The values given range from 0.167 to 0.305 for the thick portion and from 14.77 to 20.7 for the thin portion. The data were irregular. Some of the flow curves were far from linear, others indicated a neg yield value and some had a neg slope. Many expts proved that it was impossible to obtain accurate values with the material sepd by simple decantation even where the material was as fresh as possible. The effort was made to obtain consistent results by passing the material repeatedly through the pores of a Gooch crucible, but the results were unsatisfactory, since the fluidity obtained depended upon the no. of times the material was passed through the crucible. I. URBNE C. BINCHAM

The temperature dependence of the solubility of inert gases in liquids. G. TAMMANN. *Z. anorg. allgem. Chem.* 194, 159-61 (1930) — Values for the expression $(1/\log f) - (d \log f / dT)$ were obtained for He, Ne and Ar for temps between 15° and 37°. If Q , the heat tone, is negligible in comparison to A , the external work, and the gas soly, f , is less than 1, i. e., the concn of gas in the liquid is smaller than that in the gas phase, the gas soly increases with increasing temp. If Q is greater than the heat corresponding to the work, A , the soly decreases with increasing temp. If $Q=0$, $(1/\log f) (d \log f / dT)$

should be equal to $-1/T$. This is found to be only approx true for single values but for the av. of all the values for the soly. of the gases in the liquids used it is accurate. MeOH, EtOH, acetone, benzene and cyclohexane were used. The soly. of A in water decreases appreciably with increasing temp. The soly. of B is independent of the temp between 15° and 30° while that of C decreases somewhat. ALLEN S. SMITH

Solubilities in hydrogen fluoride. P. A. BOND AND V. M. STOWE. *J. Am. Chem. Soc.* 53, 30-4 (1931).—The solubilities of different salts in anhyd. HF (prepd. by distn. of dry KHF_2) were detd. in a "bomb" of Monel metal. The soly. of LiCl , expressed as mols. LiF per mol. HF, varies between 0.33 and 0.40 over the range of 0° to 40°. ZnF_2 , MgF_2 , and CaF_2 are very insol. KF reacts with liquid HF. CrF_3 is sol. to an undetd. extent. The solvent action of HF is similar to that of water. D. B.

The solubilities of alkali chlorides and sulfates in anhydrous alcohols. F. R. KIRBY AND H. L. DUNLAP. *J. Am. Chem. Soc.* 53, 391-4 (1931).—At 20° the solubilities of NaCl, KCl and Na_2SO_4 in MeOH, EtOH, PrOH, iso-PrOH, BuOH and iso-BuOH over the temp. range 20-50°, in moles times 10^3 of salt per mole of alc., are

	MeOH	EtOH	PrOH	iso-PrOH	BuOH	iso-BuOH
NaCl	0.778	0.115	0.004	0.006	0.007	0.003
KCl	0.833	0.127	0.007	0.123	0.008	0.003
Na_2SO_4	0.555	0.143	insol.	0.089	insol.	insol.

R. A. FOULF

The solubilities of silver salts. KOSAKU MASAKI. *Bull. Chem. Soc. Japan* 5, 345-8 (1930).—The soly. was detd. from the e m f. of the cell $\text{Ag} | 0.01 \text{ N AgNO}_3 | \text{NH}_4\text{NO}_3(\text{satd.}) | \text{Ag salt in } 0.1 \text{ N K salt (except KCN which was } 0.01 \text{ N)}$ by the equation $e = 0.53 \times 2 \times 0.0591 \ln(ca/x)$, where 0.53 is the transference no., a is 0.93, c is 0.1 or 0.01 and x is the concn. of Ag^+ in the K salt soln. The mean e m f. and the soly. in mol/l. at 18° are: SCN 0.549 v. and 1.0×10^{-4} , Br 0.570 v. and 7.0×10^{-7} , CN 0.571 v. and 2.1×10^{-7} ; and CO_3 0.205 v. and 1.1×10^{-6} . V. F. HARRINGTON

Theory of hydrotropy. CARL NEUBERG AND FRITZ WEINMANN. *Biochem. Z.* 229, 466-79 (1930).—The ability of many salts of org. acids to make insol. substances, especially nonelectrolytes, water-sol. is designated as hydrotropy. Expts. with K salts of campholic acid lead to the conclusion that this phenomenon is explained in the formation of a compd. between the hydrotropic salt and the dissolved substance. S. M.

Refraction of electric waves ($\lambda = 12 \text{ cm.}$) in some electrolytes. K. ZAKRZEWSKI AND T. NAYDER. *Bull. intern. acad. polonaise* 1930A, 30-41.—Measurements were made of the n of water and of aq. solns. for damped elec. waves of wave length 12 cm. The method employed consists in measuring the deviation produced in a narrow pencil of waves by a hollow glass prism filled with the liquid. The angle of the prism used was very small ($3^\circ 56'$), but large enough to eliminate inaccuracy due to reflections on the interior of the prism. The oscillator and detector were of the types described by Zakrzewski (*C. A.* 22, 2103). Within expd. error the n of water for this wave length is the same as that for very long waves. The n s of Cu sulfate of up to 3% concn. and of NaCl solns. of up to 1% show no perceptible difference from those of pure water. B. C. A.

Electrical conductivity and viscosity of aqueous solutions. M. CRANZO AND G. CLUZET. *Compt. rend. soc. biol.* 100, 1205-7 (1929).—The elec. conductivities of solns. of KCl, NaCl, NaOH, Na_2SO_4 , HCl, H_2SO_4 , and CuCl₂ are reduced by addn. of sucrose, glycerol, agar or gelatin. B. C. A.

The thermodynamic properties of molten solutions of lithium bromide in silver bromide. E. J. SALSTROM AND J. H. HILDEBRAND. *J. Am. Chem. Soc.* 52, 4650-5 (1930).—The energy changes of molten AgBr upon diln. with LiBr were detd. at mole fraction of LiBr varying from 0 to 0.89 between 430° and 600°. The free energy of formation of pure molten AgBr is given by the equation $\Delta F_a = -21,510 + 6.7 t$ between 430° and 575°. Density detns. of 0.5 mole fraction LiBr in AgBr were made between 517° and 555° and may be expressed by the equation $d = 4.504 - 0.000577 t$. The results are discussed on the basis of simple assumptions as to ionization, but without taking into account any changes in interionic forces upon diln. I. H.

A discussion of the new view in the theory of solutions. GIOVANNI B. BONINO. *Atti III congresso naz. chim. pura applicata* 1930, 231-48.—A review of the Debye-Huckel theory. J. B. AUSTIN

Dissociation theory of solutions. BERKELEY. *Nature* 126, 313 (1930).—Two formulas are given connecting the osmotic pressure, sp. vol., vol. of solute and the no. of g. mols. of solvent of salt solns. The formulas held for aq. solns. of cane sugar, α -

methylglucoside and isodulcitol at both 0° and 30° within variations of about 5%.
J. W. SHIPLEY

Proton and electron activity in solvents in general. GAROLD SCHWARZENBACH *Helv. Chim. Acta* 13, 870-96(1930)—Modern conceptions of acids and bases are reviewed critically with 81 citations. Brønsted's definitions (*C. A.* 24, 2922) are somewhat modified. Acids are substances that can yield protons, bases are substances that can take up protons. Acids and bases may be either uncharged molecules or ions. An acid can yield a proton only in the presence of a base that can take it up. There are no free protons in solns. The proton is in homopolar combination with the acid and can probably penetrate even into the electron shell. The normal acidity potential of an acid, Γ_{ac} is defined by the equation $\Pi = \Gamma_{\text{ac}} + RT/F \ln[\text{acid}]/[\text{base}]$ in which Γ is the potential of a H_2 electrode and $[\text{acid}]$ and $[\text{base}]$ are the molar concns. multiplied by the activity coeffs. for the acid and base, resp. Γ_{ac} varies with the dielec. const. of the solvent but is independent of the acidity or basicity of the solvent. Strong acids have high Γ_{ac} , weak acids low Γ_{ac} . The difference in behavior between strong and weak acids is more marked the less basic the solvent. Thus HOAc behaves much like HCl in liquid NH_3 , in H_2O the two acids differ materially, and in EtOH they differ still more markedly. In non basic solvents Π attains enormous and indefinite values because pure acids have no tendency to yield protons. Study of the acidimetry of such solns. is unprofitable. In the absence of solvent, acids have a self buffering action. Study of such systems is very difficult because traces of impurities affect the potential enormously. Proton active substances in the pure state are always both basic and acidic to some extent at least, that is, they are amphoteric.
I. L. BAOWE

The dissociation of strong electrolytes. III. Complete dissociation and optical properties. MORRIS B. JACOBS AND CECIL V. KRAV *J. Phys. Chem.* 35, 490-7(1931); *C. A.* 24, 3417, 3696—Additivity of color or light absorption, of refraction and of dispersion by the ions in solns. of strong electrolytes was not found in all cases. The most exact recent work shows discrepancies that must be explained by some means other than the degree of ionization.
H. F. JOHNSON

The electrode potentials of the silver halide electrodes against a mixture of different halide solutions. JUZABURO TANIDA *J. Biochem. (Japan)* 12, 411-8(1930)—The potential of Ag halide electrode immersed in a soln. contg. halogen ions against a calomel electrode, is determined mainly by the activity of the halogen ions of the largest mol. wt. Other halogen ions of smaller mol. wt. as well as anions other than the halides are indifferent. The potential is the same whether a AgCl , AgBr , AgI or a blank Ag electrode is used.
S. MORGULIS

The reactions between atoms and molecules of nitrogen and hydrogen. W. STEINER *Z. Elektrochem.* 36, 807-14(1930)—An app. is described for separately activating N and H by elec. discharge, allowing the gases to react and finally collecting the products of reaction. H atoms show no reaction with N mols. The reaction between N atoms and H atoms gives NH_3 , while the reaction between N atoms and H mols. furnishes N_2H_4 and NH_3 , the former in much the smaller quantity. It seems probable that the NH_3 is formed through a triple collision: $\text{N} + \text{H} + \text{X} \rightarrow \text{NH}_3 + \text{X}'$, followed by the reaction of NH_3 with mol. H : $\text{NH}_3 + \text{H}_2 \rightarrow \text{NH}_4$, while the N_2H_4 arises thus: $\text{N} + \text{H}_2 + \text{X} \rightarrow \text{N}_2\text{H}_4 + \text{X}'$ and $2\text{NH}_3 \rightarrow \text{N}_2\text{H}_4$. If such be the case, only every hundredth or thousandth triple collision between N and H_2 is effective in causing reaction.
W. C. FERNFELT

The action of carbonic acid, under pressure, on salts of the alkaline earth metals. I. Action on calcium phosphate. ERNST MÖLLER AND JULIUS KNÖFEL *Z. anorg. allgem. Chem.* 194, 258-60(1930)—Under pressure a soln. of CO_2 decomposes secondary and tertiary Ca phosphate forming $\text{Ca}(\text{HCO}_3)_2$. At 60 atm the soly. of CaHPO_4 is 6 times that at ordinary pressure. Soly. decreases with increase in temp. $\text{Ca}_3(\text{PO}_4)_2$ dissolves to the extent of 0.754 g. of P_2O_5 per l. Bone ash gives a soln. contg. 0.589 g. and 0.333 g. P_2O_5 per l. at 0° and 20°, resp. Phosphoric acid sol. to the extent of 0.035 g. P_2O_5 per l. at 0°. The action on natural phosphates is less than that on synthetic samples because of the greater compactness of the former. II. Action on alkaline earth silicates. ERNST MÖLLER AND WALTER LINDERGER *Ibid.* 201-7—At 20° and 50 atm. 2 or more hrs. action forms a soln. of bicarbonate equiv. to 3.49, 3.79 and 5.75 g./l. of CaSiO_3 , SrSiO_3 and BaSiO_3 , resp. At 0° the action is greatest, the soly. decreasing with increase in temp. The min. soly. is reached at 100°. $[\text{H}^+]$ is greatest at 30 atm. Hydrolysis and adsorption effects obscure the reaction mechanism. The soly. of ignited silicates at 50 atm. is 10 times that at ordinary pressure and $1/2$ that of non ignited samples. A method for the prepn. of carbonate free silicates is given.
B. A. SORLE

Action of water vapor on sodium chloride at higher temperatures. STANISLAW BRITZNAIDER *Roczniki Chem.* 10, 729-35 (735 German) (1930) —The equil of the reaction $\text{NaCl} + \text{H}_2\text{O} = \text{NaOH} + \text{HCl}$ was detd. The value of K_p (= $p_{\text{HCl}}/p_{\text{H}_2\text{O}}$) was found approx 1.6×10^{-4} at 750° . J. KUCERA

Hydrolysis equilibria of triphenylchloromethane and triphenylbromomethane. The free energy of some reactions of triphenylchloromethane, triphenylbromomethane and triphenylcarbinol. J. O. HALLFORD *J. Am. Chem. Soc.* 53, 105-12 (1931) —A C_6H_5 soln of the carbinol or carbinol halide was shaken in contact with an aq soln of the halogen acid of such strength that at equal measurable concns of carbinol and halide were produced. For the reaction $(\text{C}_6\text{H}_5)_3\text{CCl}$ (in C_6H_5) + $\text{H}_2\text{O}(g) = (\text{C}_6\text{H}_5)_3\text{COH}$ (in C_6H_5) + $\text{HCl}(g)$, $K_{300} = 0.0518$ and $\Delta F_{300} = 1755$ cal. For the reaction $(\text{C}_6\text{H}_5)_3\text{CCl}(s) + \text{H}_2\text{O}(g) = (\text{C}_6\text{H}_5)_3\text{COH}(s) + \text{HCl}(g)$, $K_{300} = 0.219$ and $\Delta F_{300} = 900$ cal. In obtaining these values K was detd by analysis of equil mixts and ΔF was calcd from the equation $\Delta F = -RT \ln K$. Similarly 14 other free energy relations of $(\text{C}_6\text{H}_5)_3\text{COH}$, $(\text{C}_6\text{H}_5)_3\text{CCl}$ and $(\text{C}_6\text{H}_5)_3\text{CBr}$ were detd or calcd. These include solubilities and ionizations in C_6H_5 and in H_2O as solvents. F. D. BROWN

Double decomposition in the absence of solvent. I. The reciprocal system: $\text{AgCl} + \text{KBr} \rightleftharpoons \text{AgBr} + \text{KCl}$. A. P. PALFREY *Acta Univ. Acad. Medice, Series VI, Chem.* No. 4, 3-17 (1930) —The reciprocal system $\text{AgCl} + \text{KBr} \rightleftharpoons \text{AgBr} + \text{KCl}$ was studied. The systems $\text{AgBr}-\text{AgCl}$ and $\text{KBr}-\text{KCl}$ form unbroken series of solid solns., and the systems $\text{AgBr}-\text{KBr}$ and $\text{AgCl}-\text{KCl}$ form eutectics. In order to det the space equil diagram of the quaternary system the following systems were studied and their phase diagrams detd by means of heating and cooling curves. $\text{AgCl}-\text{KBr}$, eutectic (25.1 mol percent KBr) at 320° . $\text{AgBr}-\text{KCl}$, eutectic (23.5% KCl) at 318° . $\text{AgBr}-(50\% \text{KBr} + 50\% \text{KCl})$, eutectic (74.8 AgBr, 12.6% KCl, 12.6% KBr) at 302° . $\text{KBr}-(70\% \text{AgBr} + 30\% \text{AgCl})$, eutectic (51.1% AgBr, 23.2% AgCl, 22.7% KBr) at 318° . $\text{KCl}-(35\% \text{AgBr} + 65\% \text{AgCl})$, eutectic (24.1% AgBr, 49.1% AgCl, 21.5% KCl) at 321° . $\text{KBr}-(42.5\% \text{AgCl} + 57.5\% \text{AgBr})$, eutectic at 325° . (50% AgBr + 50% AgCl)-(50% KBr + 50% KCl), eutectic (10.9% KCl, 10.9% KBr, 39.1% AgBr, 39.1% AgCl) at 323° . (62.5% AgCl + 37.5% AgBr)-KBr, eutectic (48.2% AgCl, 23.8% AgBr, 23% KBr) at 323° . The liquidus surface of the entire quaternary system consists in 2 fields of crystn intersecting along the eutectic line. The lowest point on the eutectic line (23.5°) corresponds to the eutectic compn of the system $\text{AgBr}-\text{KBr}$. Shifting of equil toward AgBr was noted, it was indicated on the open-space model by the wave on the liquidus surface, and on the projection of the eutectic line, by the max located nearer to the AgBr field. This agrees with the lower soly. of AgBr. II. The reciprocal system: $\text{AgBr} + \text{KI} \rightleftharpoons \text{AgI} + \text{KBr}$. *Ibid* 18-40 —This system resembles the previously investigated system $\text{AgCl} + \text{KBr} \rightleftharpoons \text{AgBr} + \text{KCl}$ in relation to its general character. The binary systems $\text{AgBr}-\text{AgI}$ and $\text{KBr}-\text{KI}$ both form continuous series of solid solns. The binary system $\text{AgBr}-\text{KBr}$ forms a eutectic (68 mol percent AgBr, 285°), and the system $\text{AgI}-\text{KI}$ forms an unstable compd $\text{KI} \cdot 4\text{AgI}$, and a eutectic at 213° . Several diagonal sections of the "equil square" of the quaternary system were studied by means of heating and cooling curves and the following results were obtained. The system $\text{AgBr}-\text{KI}$, eutectic (80.3 mol percent AgBr) at 264° . $\text{AgI}-\text{KBr}$, eutectic (84 mol % AgI) at 361° . $\text{AgBr}-(50\% \text{KBr} + 50\% \text{KI})$, eutectic (76.4% AgBr, 11.8% KBr and 11.8% KI) at 273° . $\text{AgI}-(50\% \text{KBr} + 50\% \text{KI})$, eutectic (61.4% AgI, 9.3% KI, and 9.3% KBr) at 336° . (70% AgI + 30% AgBr)-KI, eutectic (25.2% AgBr, 58.6 AgI, 16.2% KI) at 300° . (50% AgI + 50% AgBr)-KI, eutectic (42.3% AgI, 42.3% AgBr, 15.4% KI) at 334° . (6% AgBr + 94% AgI)-KI, eutectic (21.5% KI, 71% AgI, 4.5% AgBr) at 283° , and a transition point (21.5% KI, 73.8% AgI, 4.7% AgBr) at 285° . (7.5% AgBr + 92.5% AgI)-KI, eutectic (20.4% KI, 73.6% AgI, 6% AgBr) at 300° , and the transition point (21.3% KI, 72.4% AgI, 5.9% AgBr) at 293° . (87.5% AgBr + 12.5% AgI)-KI, eutectic (17.8% KI, 71.9% AgBr, 10.3% AgI) at 253° . (74% AgBr + 26% AgI)-KI, eutectic (16.9% KI, 61.5% AgBr, 21.6% AgI) at 253° . The system (9% AgBr + 91% AgBr)-KI, the triple transition point of simultaneous crystn. of the solid soln AgI, AgBr, the solid soln KI, KBr and the compd $\text{KI} \cdot 4\text{AgI}$, occurs at 317° , and corresponds to the compn 20.5% KI, 7.1% AgBr and 72.4% AgI. The displacement of equil. manifests itself here even more clearly than in the system $\text{AgCl} + \text{KBr} \rightleftharpoons \text{AgBr} + \text{KCl}$; the system gives in the direction of the upper arrows a pos thermal effect of 8.8 cal. The definite unstable compd $\text{KI} \cdot 4\text{AgI}$, decomposable at 267° , is formed in the lateral binary system $\text{AgI}-\text{KI}$, and forms its crystn field on the liquidus surface of the quaternary system. The latter, therefore consists of the 3 fields of crystn: (1) solid solns of KI and KBr, (2) solid solns of AgBr and AgI, and (3) compd. $\text{KI} \cdot 4\text{AgI}$.

The eutectic lines meet at the triple point (317°, 20.5% KI, 7.1% AgBr, 72.4% AgI). The eutectic point of the system KI-AgI is, however, below the triple point (213° at 20.5% AgI). The projection of the eutectic line on the vertical AgBr-AgI plane rises rapidly to a max (about 15% AgBr). The system described here as well as the system $\text{AgCl} + \text{KBr} \rightleftharpoons \text{AgBr} + \text{KCl}$ can be regarded as a representative of the type of systems that forms a transition from purely reversible to irreversible systems.

B. N. DANILOFF

Contribution to the study of the equilibrium, $\text{FeS}_2 \rightleftharpoons \text{FeS} + \text{S}$. FRANZ DE RUDNER. *Bull. soc. chim.* (4), 47, 1225-1251 (1930).—The reversibility of the reaction, $\text{FeS}_2 \rightleftharpoons \text{FeS} + \text{S}$, was established by distn of pyrites. A new synthesis was demonstrated. The intermediate compounds, FeS_2 to FeS_{12} , postulated by others to be formed by the desulfurization of pyrites, do not exist and are considered as simple crystallographic forms of mixts of FeS_2 , FeS and S . During the distn As_2S_3 , present as an impurity, is given off together with SO_2 and H_2S formed from water present in the pyrites. Dissocn. curves for pressures of 9 mm to 768 mm with extreme temps of 600° and 682°, resp., exhibit a clearly defined flat portion corresponding to the heat of absorption. The heat of reaction absorbed by the liberation of one S atom from FeS_2 is 18,513 cal., as calcd. from the Nernst equation. The roasting of pyrites in industrial furnaces gives FeS as an intermediate product, involving as a first reaction the dissociation of pyrites.

H. W. WALKER

The affinity of metals for sulfur. I. The thermal equilibrium between hydrogen and the sulfides of iron, tin, cadmium, bismuth and antimony. L. V. BAJTZEK AND A. T. KAPUSTINSKI. *Z. anorg. allgem. Chem.* 104, 323-50 (1930).—The flow method of Jellinek (*C. A.* 19, 1313, 23, 5400) was used to study the above equl. between 400° and 1000°. By combining their data for the equl. const. (in the sulfide reduction) with those in the literature on the dissociation const. of H_2S , B and K. calcd. the partial pressure, p , of S_2 over the various sulfides. Within the range studied $\log p$ is a linear function of T . At ordinary temp. the metals stand in the following order of decreasing affinity for S: Cd, Sn, Fe, Sb and Bi. The mol. heats of formation (in large cal. at 0°) from rhombic S and metal are: CdS 19.80, SnS 20.51; FeS 17.08, Bi_2S_3 42.50, H_2S_2 67.20.

OSCAR T. QUINBY

Velocity of solution of aluminum in alkali solutions. M. TRANTNER-SHIVER AND W. WITTANDT. *Bull. intern. acad. polonaise* 1930A, 50-61.—In agreement with Jablonsky, Hermanowicz and Wajschelisz (*C. A.* 23, 3390), it is found that the rate of soln. of Al in NaOH soln. is proportional to the free surface area of the metal and to the sq. root of the NaOH concn. KOH, LiOH, Ba(OH)₂ and Ca(OH)₂ solns. all show the same order of reaction as the NaOH soln. Ca(OH)₂ shows a lower velocity of reaction than the corresponding concns. of the other hydroxides, but this is probably due to the formation of a protective film of the insol. Ca aluminate. With weak bases such as NH_3 , ethylamine, diethylamine and triethylamine the reaction velocity is more closely proportional to the fourth root of the base concn., but this is proportional to the sq. root of the OH ion concn. for weak bases. It is concluded that the velocity of soln. of Al depends only on the OH ion concn. and this is supported by the fact that for different bases the rate of soln. is approx. proportional to the OH ion concn. of the solns. This yields a convenient method of detg. the hydrolysis consts. of Na_2CO_3 and K_2CO_3 , which are both about 1.4×10^{-4} . The reaction velocities have a temp. coeff. of about 2 for 10° and are independent of the rate of stirring of the soln.; hence it is a true reaction velocity which is measured and not a rate of diffusion. B. C. A.

Reactions involving hydrogen peroxide, iodine and iodate ion. I. Introduction. WM. C. BRAY AND HERMAN A. LEBELAVSKY. *J. Am. Chem. Soc.* 53, 38-44 (1931).—A crit. review and discussion of various reactions of H_2O_2 in acid solns. contg. I and iodate or iodide are based on the following table.

Reaction of H_2O_2 with

	I^-	I_2	$\text{I}_2 + \text{IO}_3^-$	IO_3^-
$10^{-10} M \text{ I}^-$ (a) —	(d) $\text{I}_2 \rightarrow \text{I}^-$ very rapid	(g) $\text{I}_2 \rightarrow \text{I}^-$ very rapid	(j) —	
$10^{-4} M \text{ I}^-$ (b) $\text{I}^- \rightarrow \text{I}_2$ moderate	(e) $\text{I}_2 \rightarrow \text{I}^-$ moderate	(h) —	(k) $\text{IO}_3^- \rightarrow \text{I}$ and I^- very slow	
$10^{-4} M \text{ I}^-$ (c) $\text{I}^- \rightarrow \text{I}_2$ faster than e	(f) —	(i) $\text{I}_2 \rightarrow \text{IO}_3^-$ rapid	(j) $\text{IO}_3^- \rightarrow \text{I}_2$ faster than k	

II. The preparation of iodic acid. Preliminary rate measurements. WM. C. BRAY AND A. L. CAULKINS. *Ibid.* 44-8.—Values of K , the first-order rate const., are nearly

const., $0.6 \approx 0.1$ at concns of I_2 from 1×10^{-3} to 3×10^{-3} *M* but always fall off markedly at lower concns. K is independent of the concn of iodate acid and peroxide within the ranges investigated. The simplest explanation is that the first-order reaction is the rate of hydrolysis of I and the that rapid decrease of K at low I concn is due to the effect of the reverse reaction as the hydrolysis equl is approached. The oxidation of I_2 by H_2O_2 furnishes a direct method of investigating the rate of hydrolysis of I .

DON BROST

The velocity of decomposition of the carbonato-tetrammine-cobaltic ion and its dependence on the hydrogen-ion concentration. KAT JULIUS PROFFER. *J Am Chem Soc* 53, 18-30(1931).—In glycolate buffer solns the reaction $[Co(NH_3)_4CO_3]^{++} + 2H^+ + H_2O \rightleftharpoons [Co(NH_3)_4(H_2O)_2]^{++} + CO_2$ is practically complete. In some acetate buffers it is balanced. "When the CO_2 pressure is low and the H ion concn not too small, the velocity may be expressed as the sum of 2 terms, one being independent of the H -ion concn, the other being proportional to the H ion concn. No general acid catalysis was found." (Cf. Brønsted, *C A* 22, 4734.) DON BROST.

Kinetics of the formation of malonamide from ethyl malonate and ammonia in homogeneous solution. A reaction of the third order. KENNETH C BAILEY. *Proc Roy Irish Acad* 39B, 567-73(1930).—Although the reaction between $CH_3(CO_2Et)_2$ and NH_3 in homogeneous medium might easily proceed stepwise through the formation of the half amide, it proves to be a true termol reaction. The reaction was studied in dil EtOH at 0° , 25° and 71.8° . The resp values for $k \times 10^3$ are 31.1, 117 and 177. The temp coeff is 1.65-1.70. As the concn of the EtOH is increased, the reaction velocity decreases, thus under the conditions of these expts the reaction velocity in 75% EtOH was less than $1/2$ its value in 44.5% EtOH. The reaction, however, continued termol. An unexpected increase in k was obtained toward the conclusion of the expt carried out at 73.8° and it is uncertain whether this was due to unrecognized exptl error or to some accelerating factor developing in the reaction.

C H PRIT

Pressure-temperature diagrams for definite phase numbers. P NIGGLI. *Chem Erde* 5, 201-24(1931).—Fifty pressure temp diagrams are given for 1 to 5 phase systems.

B C A

Theory of arranged mixed phases. CARL WAGNER AND WALTER SCHOTTKY. *Z. physik. Chem., Abt. B*, 11, 163-210(1930).—Binary systems in which the atoms or mols are shown by x ray investigations to be arranged in regular lattice arrangement are designated "arranged" mixed phases, as contrasted with those in which the atoms or mols. are distributed haphazardly. A surplus of one or the other component causes a lattice disturbance. The superfluous atoms may (1) occupy an intervening position in the space lattice, (2) produce vacancies in the space lattice, or (3) occupy a position in the space lattice of the other component. If the component in excess requires less space than the other, position (1) is noted, if a larger space, position (2) is noted. Where both components require about the same space, as in the case of most intermetallic compds, position (3) occurs. The chem potential and activity, calcd thermodynamically, det. the relation between the elec resistance and the compn. The systems, U-O, Fe-N, Hg-Zn, Au-Hg and Au-Cu are considered. CURTIS L WILSON.

Fusion diagram of the systems mercuric bromide-mercuric sulfate and mercuric chloride-mercuric sulfate. M. PAIC. *Compt rend* 191, 1337-9(1930).—By use of a method previously described (*C. A.* 24, 4204) diagrams are derived for $HgBr_2$ - $HgSO_4$. A eutectic mixt. fusing at 236° contg less than 1% $HgSO_4$ is indicated. The fusion curve mounts rapidly, 500° being the fusion point for 30% $HgSO_4$. From this point values are extrapolated up to the fusion point of $HgSO_4$. The soly of $HgSO_4$ in fused $HgCl_2$ is less than in $HgBr_2$, the fusion point of a mixt. contg 10% $HgSO_4$ is 512° .

M McMAHON

System chromium-carbon. A WESTGREN AND G PHRAGMÖN. *Z anorg allgem Chem.* 187, 401-3(1930).—The results of Krawczek and Sauerwald (*C A* 24, 1049) contradict in certain important respects those previously obtained by the authors (*C A* 22, 2913). The Cr-Cr carbide eutectic is at 3.4% C and not at 4.5%. A cubic carbide of the formula Cr_7C_3 is indicated both by x-ray and micrographic observations in addn to the carbides Cr_7C_3 and Cr_3C_2 . There is no series of mixed crystals in the concn. interval around 0% C. These discrepancies are attributed to an error on the part of K and S. in the prepn of sections for examn.

B C A

The vapor pressure and the activity of a volatile component in binary alloys at high temperatures. KARL JELLINEK AND GUSTAV A. ROSNER. *Z. physik. Chem., Abt. A*, 152, 67-94(1931).—The vapor pressure of Cd alloys with Pb, Sn and Cu at concns of 10 to 60% Cd and temps from 500° to 700° and the vapor pressure of Cu-Zn of the

concentrations from 600° to 800° were measured. The activities of Cd and Zn were calculated and compared with the electrometric measurements of Taylor. The heat of vaporization and mixing were also calculated. V. F. HARRINGTON.

The electrical conductivity and the thermal expansion of magnesium-cadmium alloys. G. GAUBE and E. SCHMIDT. *Z. anorg. allgem. Chem.* 194, 190-222 (1930). — Investigation of Mg-Cd alloys by thermal analysis does not yield max or min points on the solidification curve. An exhaustive systematic investigation of the constitution diagram of the alloys was undertaken by elec. cond. and thermal expansion methods from room temp. to near the m. p. The cond. isotherm of the alloys at 50° shows 3 decided points at 25, 50 and 75% Mg. This confirms the compd. $MgCd$ and proves the previously unknown compds. Mg_2Cd and $MgCd_2$. The 3 compds. are revealed only at lower temps.; $MgCd_2$ results at 80° on cooling the Cd-rich α mixed crystals sepd. from the melt, $MgCd$ on cooling the Mg-rich β mixed crystals to 150°. Cd and $MgCd_2$, as well as $MgCd$, and $MgCd$, form a continuous series of mixed crystals. In comparison, $MgCd$ and $MgCd_2$ are only limitedly sol. in one another. At room temp. stable Cd-rich α' mixed crystals extend from 0 to 65% Mg. Mixed crystals between the concn. boundary of the β' mixed crystals cannot be accurately detd. but lie in the neighborhood of 60% Mg. The transformation of α' - or β' -mixed crystals into the α or β , stable at high temps., results on heating through heterogeneous decomposition regions whose boundary curve is placed between 20 and 83% Mg. It was shown that in the region of 40-60% Mg and of 70-80% Mg the elec. resistance and the length of the test piece increase with heating, first with a const. temp. coeff. and then in a temp. region of 70-80° below the true transformation interval the temp. coeff. slowly increases. In the transformation interval it increases suddenly and then decreases to a small const. value. The transformation therefore proceeds in 2 steps. The first occurs with the temp. slowly increasing, disarranging the ordered distribution of the atoms, stable at low temps. The second follows the transition of the yet partially ordered into the completely irregular distribution of mixed crystals stable at higher temp. ALLEN S. SMITH.

Lattice deformation and catalytic activity. NICOLA PARRAVANO. *Atti III congresso naz. chim. pura applicata* 1930, 45-50. — The behavior of a no. of metals and metallic oxides as catalysts is considered. It is claimed that in some cases catalysis is due to the crystals of an amorphous powder. It may also be due to a disturbance of the force field around the atoms or molecules by deformation of the crystal lattice. J. B. AUSTIN.

The catalytic activity of nickel in the form of aerosol and aerogel. W. E. GIBBS and H. LIANDER. *Trans. Faraday Soc.* 26, Pt. II, 658-62 (1930). — Ni aerosol was prepd. in gas mixts. of CO and H_2 (1) by forming an elec. arc between Ni electrodes and (2) by thermal decomposition of Ni carbonyl. The aerosol prepd. by these 2 methods showed almost no catalytic activity in the formation of C_2H_4 from CO and H_2 . Ni aerogel was formed (1) by the condensation of Ni vapor formed when an arc was struck between a graphite cathode and a Ni ingot anode, (2) by the decomposition of Ni carbonyl. These aerogels were used as catalysts in the formation of C_2H_4 from CO and H_2 . They exhibited almost no activity under the same conditions that a Ni + 15% Al_2O_3 catalyst gave about 90% conversion. The Ni aerogel from the arc showed greater activity in the hydrogenation of C_2H_4 . Conclusion. Finely divided Ni prepd. by these methods did not produce active catalysts in spite of the large surface formed. L. H. REVERSON.

Catalytic experiments with a high-pressure circulation apparatus. E. BERL and R. BREMMANN. *Z. anorg. Chem.* 44, 34-9 (1931). — Water gas was subjected to catalytic high-pressure synthesis in a circulation app. in which the liquid reaction product is sepd. from the gas mixt. A sketch of the app. and details of the design of the high-pressure bomb, sepd. vessel, fittings and device for measuring gas velocity are given. The results in general, confirmed data reported in the literature. Alkali-free ZnO catalyst forms MeOH chiefly. In contrast, higher alcs., acids and aldehydes are formed by the use of an alkali-contg. ZnO. S has an injurious effect on the catalyst. The prepn. and reduction of the catalyst are of great importance. The gas velocity over the catalyst has considerable influence on the reaction. With increasing velocity more unstable intermediate compds. are obtained, while at low velocities the stable end products predominate. ALLEN S. SMITH.

The catalytic oxidation of carbon monoxide. J. C. W. FRATER. *J. Phys. Chem.* 35, 405-11 (1931). — Carefully purified, finely divided MnO_2 , Co_2O_3 , and Ni_2O_3 are extremely active catalysts for the union of CO and O_2 at temps. as low as 0°. Extreme purity, not the presence of promoters, is important for catalytic activity. Impurities may be removed by electrolysis of an aq. suspension of the oxides. None of 17 other

unnamed oxides studied was effective. The 3 active oxides differ from the others in one significant particular, namely their composition is indefinite. By an adsorption apparatus it was shown that the O_2 pressure in eq. 1 with a sample of one of these oxides depends on the composition of the oxide. Condensation of water vapor within the pores of these oxides destroys their catalytic activity. Extreme drying does not destroy the activity. These oxides, however, completely protect a worker when used in a gas mask, but they are not yet adapted to the treatment of the exhaust gases of internal-combustion engines. References are made to several dissertations not indexed in chem. literature. I. F. BROWN.

Colloidal ferric oxide and various factors that influence its ability to catalyze the decomposition of hydrogen peroxide. I. The temperature coefficient, the effect of catalyst concentration and the effect of electrolytes. RAYMOND J. KEEFER and JAMES H. WALTON. *J. Phys. Chem.* 35, 557-77 (1931).—Decomposition of H_2O_2 was catalyzed at 50°, in the dark, by chloride-free Fe_2O_3 prep'd in a special high temp. dialyzer according to Sorum. The reaction seemed to be bimolecular. The velocity constants increased as the reaction progressed. The temp. coeff. of the reaction velocity between 30° and 60° was 2.33 for each 10° rise in temp. This indicates that a chemical change, not a diffusion, governed the rate of decomposition. The rate of decomposition was multiplied by 2.2 when the concn. of catalyst was doubled between the limits 1.26 and 2.51 g. of Fe per l. When $NaCl$, $BaCl_2$ and K_2SO_4 were added, the rate of decomposition of H_2O_2 followed typical flocculation curves which show the effects of adsorption of electrolyte on the colloidal flocculation of the sol and subsequent adsorption on the ppt'd Fe_2O_3 . When NaH_2PO_4 , $NaHPO_4$, HCl , $NaOH$ and $CuSO_4$ were added, similar curves were obtained, but these latter curves were modified by the effect on decomposition of H_2O_2 exerted by the added electrolyte. The activity of this colloidal Fe_2O_3 was about 100 times the activity of the $FeCl_3$ that it contains. The major catalytic effect is due to Fe_2O_3 and not to $FeCl_3$ as sometimes claimed. It is probable that the formation of higher oxidation products of Fe , such as FeO_2 , Fe_3O_4 , FeO or H_2FeO_4 is one of the steps in the decomposition of H_2O_2 in the presence of Fe_2O_3 . F. F. BROWN.

The heat of dilution of salts of several bivalent metals with univalent anions at high dilution at 25°. I. $MgCl_2$, $CaCl_2$, $SrCl_2$, $BaCl_2$ and $MgBr_2$, $CaBr_2$, $SrBr_2$, $BaBr_2$. F. LANGE and H. STRECKER. *Z. physik. Chem., Abt. A*, 152, 1-23 (1931).—The heats of dilution for these salts are measured at concns. from 0.1 to 0.001 M . The interpolated values for l_m and m are, at least below 0.01 M , in good agreement. The l_m curve is only approx. proportional to \sqrt{m} . The extrapolated initial slope agrees with that calculated by the Debye-Hückel theory. The deviation of m is discussed with respect to the theory, the activity and the osmotic coeffs. and the radius of the unhydrated ions. V. F. HARRINGTON.

A comparison of certain suspending agents (BARNES) 17. The reversibility of coupled reactions in biological systems and the second law of thermodynamics (BURK) 11A. The pH stability regions of the proteins (SVENBERG) 11A. The state of water in colloidal and living systems (GORTNER) 11A. Packing of atoms in alloys (WESTGREN, ALMÉN) 9.

Chemisch Jaarboekje der Nederlandsche Vereeniging. DL. II. 15th ed. Amsterdam: D. B. Centen. 256 pp. Bound, fl. 5.

Ergebnisse der exakten Naturwissenschaften, Band IX. Contg. indexes for Bände I-IX. Berlin: J. Springer. 348 pp. M. 20, bound, M. 29.40.

GRIENER, HIPPOLYTE W.: The Story of Chemistry. Volume VII of Popular Science Library. New York: P. F. Collier & Son Co. 408 pp.

GÄRTNER, PAUL: Die Welt des unendlich Kleinen. Heft 2 of "Naturwissenschaftliche Zeitschrift." Godesberg bei Bonn. Naturwissenschaftlicher Verlag. 32 pp.

HOLMES, HARRY N.: Introductory College Chemistry. Revised ed. New York: The Macmillan Co.

HUNTER, GEORGE W., and WHITMAN, WALTER G.: Teachers' Manual and Key for Problems in General Science. New York: Am. Book Co. 249 pp.

International Critical Tables of Numerical Data, Physics, Chemistry and Technology. Vol. VII (concluding volume). E. W. WASHBURN, Editor-in-Chief. Prep'd under auspices of Internatl. Research Council and the Natl. Acad. of Sciences by the Natl. Research Council of U. S. New York: McGraw-Hill Book Co. \$84 for set of 7 volumes. Reviewed in *Ind. Eng. Chem.* 23, 343 (1931). Cf. *C. A.* 24, 546.

KLOPF, WILLIAM J. The Relative Merits of Three Methods of Teaching General Science in the High School. Chicago: The Central Assoc. of Science and Mathematics Teachers, Inc. 82 pp.

KLEBER, HARALD V. Das Vorkommen der chemischen Elemente im Kosmos. Leipzig: J. A. Barth. 170 pp. M 18, bound, M 20.

DALMÁU, JOSÉ M. Pla. Elementos de química general; Físico-química. Descriptiva Industrial. Experiencias prácticas. Iniciación al análisis. Gerona: Dalmau Carles, Pla, s. a. 416 pp.

3—SUBATOMIC PHENOMENA AND RADIOCHEMISTRY

W. ALBERT NOYES, JR.

Connection between the quantum mechanics "uncertainty" and the structure of elementary particles and a calculation of the masses of the proton and electron based thereon. R. FECHT. *Z. Physik* 57, 429-46(1929); *Physik Z.* 30, 895-8(1929); cf. *C. A.* 23, 5405, 24, 1019—Theoretical. While the radius of the electron has a definite meaning in the classical electron theory, matters become very complicated on applying quantum mechanics. The Heisenberg "uncertainty" relationship is applied first to a light quantum, and then to protons and electrons. The change of matter into radiation and inversely is considered, and the masses of the electron, proton and neutron are obtained. Eddington's theory of the elementary quantum, in which it is described in 10-dimensional space, is also considered, and the above calculations are carried out.

B C A

Is there dualism between corpuscles and waves? E. SEVIN. *Ann. soc. sci. Bruxelles* 50, Ser. II, 121-5, 129-34(1930)—S rejects the explanation of electron diffraction based on the wave mechanics. He assumes (1) that when a beam of electrons of homogeneous velocity v impinges on a crystal the shock produces a "classical" wave, $s = v$, a beam of parallel rays of wave length $\lambda = h/v = (e^2/c^2)1/s$. The direction of the x ray beam is the same as if the incident electron stream were replaced by a beam of x rays. (2) This "classical" wave immediately entrains the electrons which have just been stopped and only those. The mechanism of the entrainment is identified with the Compton effect. Reply. C. MANNERBACK. *Ibid.* 125-0.

J U Austrv

Spectroscopic apparatus in industry. CHARLES C. NICHOLS. *Mech. Eng.* 53, 123-6(1931).

E 11

Canal-ray collisions. Cirk GERTSEN. *Physik Z.* 31, 918-53(1930)—The intensity of canal rays can be exactly measured by detg. the ionization in a gas chamber sealed from the evacuated app. by means of a thin window. Primary currents of 10^{-12} to 10^{-14} amp. can still be measured. The laws of canal ray collisions are well understood as those of related rays. Several briefly reviewed papers show this.

F U

The action of Geiger counting chambers. C. BOSCH AND H. KLUMM. *Naturwissenschaften* 18, 1098(1930)—An investigation was made of the influence of the surface properties of sensitive "points" as used in Geiger counters on disruptive discharges (counter impulses). Geiger Muller counters used were of the conventional type, cylindrical anode, coaxial filament. The filament could be cleaned by glowing or by ion bombardment. Filaments of Cu, Ag, Au, Fe and Zn were used in He with admixtures of H₂ or O₂. Electronegative admixtures favor the counting effect, purest electrodes and He filling of the tube gave no counting effect whatever, although recovery was noticed after some time. Immersion in liquid air had the same effect of reducing the counting effect. It is concluded that an adsorbed film on the filament has great influence on the counting efficiency, probably a H₂O film.

B J C v o H

Characteristic velocity of electrons diffused from metallic surfaces. G. BERNARDINI. *Atti accad. Lincei* 11, 1096-9(1930)—There are 3 kinds of secondary electronic emissions: reflected which do not lose any appreciable velocity; diffused, which lose a little in velocity; secondary electrons, released from the metal reflector and having considerably lower velocity. There are no characteristic velocities to distinguish these classes of electrons, but it has been found recently that at 300-500 v (Rudberg, *C. A.* 24, 3431) and also at 150 v (Brown and Whiddington, *C. A.* 21, 2221), have noted discontinuities in spectra from secondary emission suggesting inelastic collisions. B. working at 34-40 v and using Cu, Al and Zn reflectors has noted discontinuities with Cu and Zn only. No explanations have been advanced for the phenomena.

A. W. CONTIERI

Electric conduction in metals. W. M. HICKS *Nature* 126, 951-5(1930); cf. Fowler, *C. A.* 24, 3124—The fact that free electrons in metals do not affect the σ heat relation as a perfect gas is explainable by their being in the interatomic spaces an insufficient length of time to absorb the energy of the space completely. As they fall into atoms the energy is absorbed by the nucleus, the electrons are distributed in their selected orbits and the energy radiated according to the $nh\nu$ formula. This equalization of energy by the nucleus may explain the behavior of an atom as a perfect sphere and a perfect machine.

GRIGG M. EVANS

Evidence of protons in metals. A. CORN *Z. Elektrochem.* 35, 676-80(1929), cf. *C. A.* 22, 1901—Three methods are described to show that electrolytically generated H diffuses through a Pd cathode and thereby imparts to the various points on the electrode the potential which would be assumed by a H electrode. In the first expt 0.1 N H_2SO_4 soln. was electrolyzed between a Pt and a bent Pd cathode, and after stopping electrolysis the potential at different points on the latter was compared with that of the calomel electrode. At first certain points on the wire showed the O electrode potential as a result of exposure to air, but after 80 hrs. the potential became equal to that of the H electrode. The 2nd method was to observe the variation in elec. resistance of the Pd wire electrode as a result of the diffusion through it of H. In the 3rd method advantage was taken of the ready oxidation of H contained in Pd by O to form H_2O_2 .

B. C. A.

The mechanism of spark discharges. J. SLIFIAN *J. Franklin Inst.* 210, 473-5(1930)—Reply to Loeb (*C. A.* 24, 6611).

M. FARNSWORTH

Infra-red sensitive cells. FRITZ MICHELSEN *Z. tech. Physik* 11, 511-5(1930).—The action of photoelec. cells sensitive in the infra red was studied. Cells of Se-Te were prepd. by cathodic evapn. on glass, the film of alloy being extremely thin, the electrical resistance was reduced by prepng. the glass with 2 sets of dovetailed comb-shaped gratings, rubbed in with colloidal graphite and using these as electrodes. The alloy was subsequently deposited on the grating. The distance between the electrode grooves was 0.1 mm., their length 10 mm. The evapn. took place in an A glow discharge. The dark space reached from the Se-Te cathode halfway to the glass plate by proper adjustment of the A pressure. For proper grain size the plate was heated later to 210° for 2 hrs., then enclosed in an evacuated vessel. The light used for the tests on the sensitivity was filtered through proper glass filters which cut off light below 0.745 or 0.675 μ . Curves on spectral sensitivity distribution were also made. The data show an active infra-red interval from 0.75 μ to 1.2 μ for 7 to 13.6 at % Te in Se. The opinion of Gripenberg (*C. A.* 6, 449, 7, 2007) was confirmed on the sensitivity shift toward short λ of plain Se cells when used in a thin film. It is believed that the increased sensitivity of the Se-Te mixt. is due to a changed crystal form of the Se. The Case "thallofide" cell was examd. (*C. A.* 14, 2107). These cells were prepd. by thermal evapn. in O_2 at 0.8 mm. The results were quite satisfactory for practical application of the cells to fog telegraphy.

B. J. C. VAN DER HORVEN

A new kind of photoelectric cell. B. LANGE *Physik Z.* 31, 964-9(1930).—The new Cu_2O photoelec. cells and their use in photometric measurements are described. Sunlight falling on a cell of this type having a surface of 49 sq. cm. drove a small motor in the Siemens lab.

FRANK URBAN

Remark on the experiment of Rupp on a relation between electron scattering maxima and the emission of soft x-radiation. H. W. B. SKINNER *Naturwissenschaften* 18, 1097-8(1930).—Recent work of Rupp (*C. A.* 25, 870) is explainable in a different manner without relating the emission of soft x-rays to crystal structure. Whatever angle of incidence causes a max. reflection of electrons (1 e., a Bragg interference max.) will also cause a min. loss of energy of the electrons to the crystal and thus a min. chance for excitation of the x-rays. Max. x-ray intensity will, therefore, automatically coincide with max. energy scattering of the electrons. Reply: E. RUPP *Ibid.* 1098.

B. J. C. VAN DER HORVEN

Determination of atom distances in gas molecules by x-rays and cathode rays. J. M. BIJVOET AND H. J. VERWEEL *Chem. Weekblad* 27, 648-50(1930).—A review.

B. J. C. v. d. H.

Scattering of electrons by crystals and adsorbed gas films. H. J. VERWEEL AND J. M. BIJVOET *Chem. Weekblad* 27, 626-9(1930).—A review.

B. J. C. v. d. H.

Energy distribution in continuous Röntgen spectra. K. K. AGLINTZEV, *Zhur. Prikladnoi Fiziki* 6, No. 2, 38-46(1929) (Published in *Trans. State Phys.-Tech. Lab.* (Moscow), No. 11(1929)).—A general method for the theoretical detn. of energy distribution is developed analytically (from the current-voltage time relationship) and graphically (from oscillographs), and applied to inductors and transformers. It is

shown that the exptl. results after the introduction of necessary corrections, are in agreement with Kulenkampff's (*C A* 17, 1722) law of distribution of energy in Röntgen spectra

X-ray investigation of easily deformable crystals. L. CHROBAK. *Bull. intern. acad. sciences* 1929A, 497-505.—The technique of obtaining undeformed specimens of easily deformed crystals for x ray investigations is given. New app. for the work is also described including a simple precision chamber for Laue photographs

Technic of radiography by γ -rays. CHARLES S. BARRETT, ROY A. GEZELIUS AND ROBERT F. MICH. *Metals and Alloys* 1, 872-9(1930).—A technic for taking radiographs by means of γ rays is given with charts and tables of exposure times for Ra and for Ra emanation. Factors influencing the quality of radiographs and the rapidity of photographic action with γ rays are discussed

The scattering of hard γ -rays. L. LANDAU. *Naturwissenschaften* 18, 1112(1930).—It is claimed that Beck (*C A* 25, 871) is in error and that the possible scattering of hard γ rays by at. nuclei is not more than 0.0001 as great as the Compton effect

Theoretical treatment of atomic destruction. GUIDO BECK. *Physik Z.* 31, 945-6(1930), cf *C A* 25, 248.—An interpretation is given of the problem of collisions between α particles and nuclei in terms of the quantum theory of aperiodic phenomena.

Internal conversion of nuclear energy. H. CASPARI. *Nature* 126, 933-4(1930).—When a nucleus passes from an excited state to one of less energy, either a γ -quantum is given off or an extra-nuclear electron is emitted with energy equal to the γ -quantum less the ionization energy. This latter is called internal conversion of nuclear energy. The probability of this conversion will be measured by the no. of γ -quanta times a factor K (a function of the frequency). An expression for K is given, based on Dirac's equation. For hard γ rays such as those from Ra C the electron ejection is mainly nuclear

Can lead be rendered radioactive? G. GUÉZEN. *Ann. soc. sci. Bruxelles* 50, Ser. B, 117-21(1930).—An attempt was made to detect radioactivity in a series of Pb disks which had been exposed to sunlight, ultra violet light, x rays and γ rays by comparing them with a series of Al disks similarly treated. No activity was observed. Conclusion: Pb is not radioactive and the activity reported by others is probably due to a deposit from the atm.

The absolute intensities and internal conversion coefficients of the γ -rays of radium B and radium C. C. D. ELLIS AND G. H. ASTON. *Proc. Roy. Soc. (London)* A129, 169-207(1930).—The relative intensities of the γ rays of Ra B and Ra C and of the photoelec. groups assoc. with them are investigated. In a fraction, α , of the cases the radiation is absorbed in the electronic structure and gives rise to a photoelectron, in the remaining fraction, $1-\alpha$, the γ ray is emitted clear of the atom. The object of this paper is to describe measurements of the various quantities, α , and on the basis of these results to discuss the question of the mode of energy interchange between the nucleus and electron system.

The rate of loss of energy by β -particles in passing through matter. E. J. WILLIAMS. *Proc. Roy. Soc. (London)* A130, 319-27(1931), cf *C A* 24, 5212.—Exptl. results for the rate of loss of energy by β particles with velocities from 0.1-0.7 c are given. The values for slow electrons are obtained by the cloud method and those for fast electrons in expts with thin metallic foil. The results have been cor. for the effects of scattering and straggling. The observed values for the rate of loss of energy are 50-80% greater than the theoretical values as calcd. by Bohr and Gaunt using, resp., the old and new quantum mechanics.

The loss of energy by β -particles and its distribution between different kinds of collisions. E. J. WILLIAMS. *Proc. Roy. Soc. (London)* A130, 328-46(1931), cf preceding abstract.—Classical and quantum mechanics indicate approx. the same rate of loss of energy by β -particles but the distribution of the energy loss between different kinds of collisions is very different. The results of the quantum theory calcn. agree more closely with exptl., particularly in connection with the phenomenon of primary ionization. Quant. agreement is not good for the rate of loss of energy and primary ionization. For collisions in which the energy loss is large and in case of slow electrons, the agreement is within exptl. error. Gaunt's calcn. extended to the case of fast electrons and for non-relativistic quantum theory give a much smaller no. of collisions than expt. This indicates the nature of the relativity correction.

Speculations concerning the α -, β - and γ -rays of radium B, C, C'. I. A revised

theory of the internal absorption coefficient. R. H. FOWLER *Proc. Roy. Soc. (London)* A129, 1-24(1930) —This paper is a math. development of the absorption of γ -rays in a species of photoelectric effect by the planetary electrons of the parent atom, which gives rise to the sharp lines of the β ray spectrum. MARIE FARNSWORTH

A new method of analysis of groups of α -rays. I. α -Rays from radium C, thorium C and actinium C. ERNEST RUTHERFORD, F. A. B. WARD AND C. E. WYNN-WILLIAMS. *Proc. Roy. Soc. (London)* A129, 211-31(1930) —The paper describes some of the applications of the "Greiner" type of counter for α -rays, in which the ionization current due to a particle is amplified linearly by means of thermionic valves, there being no ionization by collision. By using a double ionization chamber with this type of counter it is possible, in working with complex beams of α particles, to record only those particles which stop within the chamber itself, and not those which pass right across it, thus giving a measure of the no. of particles having ranges between x mm. and $x + 2$ mm. This differential counter is applied to the study of groups of α particles emitted from various sources. The 8.6-cm. α particles from Th C', the 7.0-cm. particles from Ra C', and the 3.9-cm. particles from Po all appear to be homogeneous groups. The 5.5-cm. particles from Act C are found to consist of 2 well marked groups, differing in range by 4.2 cm. The 4.8-cm. particles from Th C form also a complex group, in agreement with the results of Rosenblum, who showed that they consisted of at least 5 groups of particles. The missing short range α particles emitted in the dual disintegration of Ra C are found and their constants determined. It is found that these α particles are present to an extent of about 1 in 4,000 of the main 7.0-cm. particles from Ra C', they appear to consist of 2 groups—the main one of range 4.1 cm. and a smaller one of range 3.9 cm. MARIE FARNSWORTH

α -Particles of the actinium series. PAOLO MISCIATTI *Atti III congresso naz. chim. pura applicata* 1930, 353-01 —M. has studied the range of α -particles emitted by radioactinium and actinium X as a function of the time elapsed since their preparation. The results are interpreted to show the appearance of various decomposition products. J. B. AUSTIN

An assumed transformation of lead. Remarks on the notes of Smits, Boutaric and Lepape. STEPHANIE MARCINEANU. *Bull. sect. sci. acad. roumaine* 13, 55-8(1930); cf. 24, 294, 2950, 3427, 4451, 4691 —S., B. and L. confirm the radioactivity of ancient Pb coverings, but consider the cause of the radioactivity debatable. M. presents arguments in favor of solar action. H. W. WALKER

Experiments to discover a stable isotope of polonium. G. V. HERVEY AND A. GREYHER. *Z. anorg. allgem. Chem.* 194, 162-78(1930) —Of the elements of the radioactive decomposition series, only elements 81, 82 and 83 have stable, as well as radioactive, isotopes. In order to discover a possible stable Po isotope, large quantities of the Te minerals, hessite, calaverite, nagayte, and the Bi minerals, Bi telluride, bismuthinite, and metallic Bi were used with Po as an indicator. The Po was added to the soln. of the mineral and at the end of the expt. was again added. A possible present stable Po isotope would separate with the Po. It is pointed out that the Po ppt. weighed only about 0.1 mg. and contained less than 1% of the element sought as determined by x-ray spectroscopic investigation. It is concluded that the minerals investigated did not contain the element sought, or contained less than 10^{-11} g. per g. of mineral. Methods were worked out for the chemical and electrolytic separation of Po from large amounts of foreign substances, especially Te and Bi. A systematic investigation of the electrolytic separation of the different metals gave the most favorable results with the use of Mo as the cathode. ALLEN S. SMITH

The isotopic constitution and atomic weights of zinc, tin, thallium and molybdenum. F. W. ASTON. *Proc. Roy. Soc. (London)* A130, 302-10(1931). —Expts. with Cd and Ge were unsuccessful. The following figures give the mass no. and percentage abundance of the isotopes, the isotopic moment, mean mass no., at. wt. and packing fraction, resp.: Zn: 64, 48.0, 65, 25, 66, 23.9, 67, 5.3, 68, 17.1, 69, 0.85, 70, 0.38, 1.41; 63.452, 63.380 \pm 0.02, -0.9. Sn: 112, 1.07; 114, 0.74, 115, 0.44; 116, 14.19, 117, 9.81; 118, 21.48, 119, 11.02, 120, 27.04; 121, 2.96, 122, 5.03; 124, 6.19, 1.77; 118.82, 118.72 \pm 0.03, -7.3. Cr: 50, 4.9, 52, 81.6, 53, 10.4; 54, 3.1; 0.31; 52.068, 52.011 \pm 0.006, -10 \pm 3. Mo: 92, 14.2, 94, 10.0; 95, 15.5, 96, 17.8, 97, 9.6, 98, 23.0; 100, 9.8, 187; 96.023, 95.97 \pm 0.06, -5.5. These results include correction of the order of abundance for Zn and Sn and the discovery of 3 new isotopes of Cr and 7 of Mo. G. M. MURPHY

The radioactive properties of rocks, soils, crude oil and waters from Southern California. J. LLOYD BOHRV. *J. Franklin Inst.* 210, 461-72(1930). cf. C. A. 24, 4699 —This paper contains measurements on waters from Lake Arrowhead, Arrowhead

Hot Springs, Harlem Hot Springs, the Pacific Ocean and from numerous wells and tunnels of the Pasadena and neighboring water supplies. It also contains measurements on rocks, soils and crude oil.

Nature and size of the luminescent center. J. EWLES. *Proc. Roy. Soc. (London)* A129, 509-19(1930).—By reasoning in accord with recent knowledge of luminescent solid solns., and free from the doubtful assumptions used by Merritt, an equation relating brightness of luminescence and concn of activator has been derived. This expression is similar to that of Brininghaus. It was found to apply to solid solns Sb_2O_3 , B_2O_3 , and PbO each in CaO , to uranium and Rhodamine 6G in water, and to fluorescein in $EtOH$.

The scattering of slow α -particles by helium. P. M. S. BLACKPITY AND P. C. CHAMPTON. *Proc. Roy. Soc. (London)* A130, 380-8(1931).—The scattering of particles by an inverse-square law is the same in classical and quantum mechanics. Mott (*C. A.* 24, 2040) by means of wave mechanics has shown the scattering to be quite different from the classical when the scattered and scattering particles are of the same kind. The scattering of α -particles by He is a case of collision of particles of the same kind that have no spin and obey the Einstein Bose statistics. Photographs of 50 000 α -rays from Po were made with the Wilson cloud method in He O_2 mixt. The results agree with Mott's theory.

Recoil atoms in a gaseous medium. LA. GOLDSTEIN. *Compt. rend.* 191, 1450-2 (1930).—It has been shown previously that the degree of activation of a negatively charged electrode immersed in a mixt. of radon and air varies with the pressure of the mixt. At low pressures the degree of activation is very feeble, because the high velocities of the recoil atoms, Ra A and Ra B, are not slowed down by the low gas pressure sufficiently for them to be diverted to the charged electrode. As the pressure increases the degree of activation of the electrode increases to a max. because the path of the recoil atoms is shortened by the increased gas pressure. Then, after passing through a max., the degree of activation of the electrode decreases as the pressure of the gas mixt. increases. In the present paper expts. are described in support of the following hypothesis regarding this decrease. As the pressure increases, the ionization of the gas increases in consequence of the better utilization of the paths of the α -rays, and the decrease in degree of activation is due to the recombination of the positively charged recoil atoms with the negatively charged ions of the gas. It is further shown that the degree of activation of an anode increases constantly with the pressure of the gas, under the same conditions that cause the degree of activation of a cathode to pass through a max. and then decrease. Also, that increase of the ionic density does not sensibly decrease the degree of activation of an anode. Expts. designed to explain this behavior of the anode are in progress.

Space distribution of H-particles liberated from aluminum by α -rays of polonium. H. POSE. *Physik. Z.* 31, 947-5(1930).—Absorption curves of H-particles obtained from Al by Po α -rays were detd. for the angles of 48° , 65° , 105° and 135° between primary and secondary rays. Excellent agreement was found with values calcd. by the method of Bothe.

Absolute intensity of the aurora line in the night sky and the number of atomic transitions required to maintain it. LORD RAYLEIGH. *Proc. Roy. Soc. (London)* A129, 458-67(1930).—The standard instrument used in measuring the green auroral light of the night sky was calibrated. If the transmitted light is assumed to consist entirely of $\lambda 5577$, the brightness corresponding to the zero reading of the instrument is found to be 8.65×10^{-4} candles/sq m. Deducting the fraction attributable to the continuous background, the brightness of the line becomes 3.2×10^{-4} candles/sq m. The energy required to maintain it is 6.4 ergs/sec./sq m. and the no. of at. transitions required to supply the energy is 1.81×10^4 sec./sq m.

Change in electron coupling in the rare gases. C. J. BAKKER. *Nature* 126, 955, *Naturwissenschaften* 18, 1100(1930).— g values for SL and JJ couplings for P_1 and P_2 levels are given for Ne I, Ar I, Kr I and Xe I. The agreement is good except in Xe I where first-order calcs. are probably not sufficiently accurate for large values of spin orbit reaction.

Interferometric measurements in the arc spectrum of iron. C. V. JACKSON. *Proc. Roy. Soc. (London)* A130, 395-410(1931).—Interferometric measurements were made of 10 lines in the Fe arc spectrum between 4000 and 4400 Å. U., by comparison with the red Cd line and the secondary Ne standards. Measurements have also been made of 68 lines between 2700 and 3100 Å. U. in the Fe arc and 5 lines in Hg from 2652-3125 Å. U. The results are in good agreement with other investigators and with calcd. results obtained from energy levels.

Apparatus for determining optical and magnetic rotatory dispersion in the ultra-violet. AUGUST HAGENRACH *Helv. Phys. Acta* 3, 168-70 (1930) UGON BRETSCHNE

The effect of crossed electric and magnetic fields on the Balmer lines of hydrogen. WALTER STELBERG *Naturwissenschaften* 18, 1098-9 (1930) —Preliminary results are given. The effect was not the appearance of new lines but consisted of a shift of the Stark-effect components and intensity changes. Weak magnetic fields do not influence the intensity distribution materially. Strong fields suppress all Stark-effect components.

B. J. C. VAN DER HOEVEN

Bergmann series in the argon spectrum. FRANK RASMUSSEN *Naturwissenschaften* 18, 1112-3 (1930). Until now Bergmann series were only known for Xe (Meggers, de Bruin, Humphreys, *C. A.* 24, 1257). Similar series have now been found for Ar by using a plane grating and neocyanine plates and an A tube at 2 mm. pressure. The results are given in tables, including derived values for the terms $3d_1$ to $3d_4$ and $2s_1$ to $2s_4$ and $3s$.

B. J. C. VAN DER HOEVEN

The fluorescence of excited mercury atoms. Z. ZAJAC *Compt. rend.* 191, 1501-6 (1930). —Hg vapor in a sealed tube was simultaneously subjected to an oscillating electric discharge and to an intense low pressure Hg vapor arc. External electrodes were used. A greenish blue luminous color was observed in the post column limited by the contour of the exciting rays. It started at a pressure of 0.01 mm. and attained its max. intensity at 0.3 mm. and at a temp. of 140° . At 1 mm. a continuous intense spectrum excited by the discharges was observed which masked the fluorescent light so that it was no longer visible at pressures above 1.5 mm. Raising the temp. of the vapor at const. pressure does not noticeably influence the phenomenon. Photographing the spectrum gives the 5461, 4358 and 4017 Hg lines which are intensified by the simultaneous excitation. Each of the 3 lines is capable of causing the fluorescence, but with different degrees of intensity. The yellow lines 5790 and 5780 do not cause it. The intensities of all lines seen during the fluorescence are proportional to the intensity of the exciting light.

I. W. LAIRD

Analysis of the ultra-violet hydrochloric acid bands. M. KULP *Physik. Z.* 21, 630-60 (1930), cf. *C. A.* 24, 5228. —The bands between 3000 Å U. and 4000 Å U. have been rephotographed and analyzed.

FRANK URBAN

New band group in the ultra-violet absorption spectrum of selenium vapor. (MILLER.) M. MORACZEWSKA. *Bull. intern. acad. polonaise*, 1930A, 17-9. —A new band group has been observed in the absorption spectrum of Se vapor in the spectral region 2100-2300 Å U. Like other Se bands they are degraded toward the red, but unlike the visible band series no rotational fine structure could be detected with the dispersion employed. The frequencies of 12 band heads have been measured and are expressed by the formula $\nu = 43,210 + 368n - 0.2n^2$. Since the bands appear in the satd. Se vapor at relatively low temp., 200° ($p = 0.001$ mm.) to 400° ($p = 4$ mm.), and since the mean frequency difference is 363 cm^{-1} as against $200\text{--}250 \text{ cm}^{-1}$ for the bands in the visible region, it is suggested either that the carrier of this spectrum is not Se_2 but a polyat. Se mol., or, more probably, that this spectrum corresponds with a different electron jump from that corresponding with the bands previously observed.

B. C. A.

Spectroscopic observations of photochemical reactions. R. MECKE *Z. wiss. Phot. (Schaum Festschrift)* 29, 72-6 (1930). —Photochem. reactions which are in agreement with Einstein's equivalence law can be regarded as single reactions. Those reactions which do not follow the equivalence law may be regarded as being chain reactions.

A. P. H. TRIVELLI

The chemical effects of cathode rays. I. The decomposition of ammonia. G. R. GEDYE AND T. E. ALLIBONE *Proc. Roy. Soc. (London)* A130, 316-66 (1931). —A beam of electrons from a Lenard tube with 210 kv. and currents up to 4 microamp. was passed through Al foil into NH_3 . The decompn. products are N_2 , H_2 , and H_2 . From 5.75 to 71.4 cm. initial gas pressure, the rate of formation of permanent gas is approx. proportional to the pressure and independent of the nature of the surface. The relative yield of N_2 is greater at low pressures for the same relative amt. of decompn. A calorimeter is described for the measurement of the energy of the electrons entering the reaction vessel. It gives results for the amt. of energy reacting per mol. of the same order of precision as may be obtained with α particles. The results are in good agreement with those obtained in the same reaction with α -particles and agree qualitatively with decompn. in the silent discharge and in ultra-violet light. For each ion pair formed, 1.2 mols. of NH_3 are decomposed.

G. M. MURPHY

Sorption and chemical reactions in atomic rays. M. LANCUCKI. *Bull. intern. acad. polonaise* 1930A, 20-9. —A study has been made of the close relationship which

exists between the absorption of N in an elec. discharge tube, Fe or Al electrodes being used, and the disintegration of these metals. The amts. of N absorbed and of Fe disintegrated are equatomic and depend on the gas pressure, the cathode fall, the c. d., and the temp. of the cathode. If the cathode is constructed of Fe nitride, the latter is disintegrated without absorption of N. The Fe nitride formed by this process is not ferromagnetic. B. C. A.

The action of electrical discharge on gaseous hydrocarbons. The effect of high-speed cathode rays on paraffin hydrocarbons. S. C. LIPP, R. M. MARKS AND G. GLOCKLER. *Trans. Am. Electrochem. Soc.* 59 (preprint) 8 pp (1931). —It has been found that the hydrocarbons methane, ethane, propane and butane are decomposed by the high speed electrons (125,000 v.) which are emitted through the thin glass window of a Slack Lenard cathode-ray tube. H₂ is one of the products of decompn. An estimate is made of the no. of mols. reacting per electron crossing the reaction tube and of the no. of ion pairs produced in the reacting gas per electron traversing it. The comparison of these ratios shows that the no. of mols. reacting and the no. of ions pairs formed are of the same order of magnitude. C. G. F.

Physical methods in the chemical laboratory. XVI. The effect of electrical discharges on gaseous elements and compounds. H. GEIGER. *Z. phys. Chem.* 44, 70-4 (1931). —The formation of active N and H, the synthesis of acetylene, the polymerization of benzene, and the decompn. of thiocyanic acid and carbonic acid in the elec. discharge are reviewed. Methods and app. are discussed. Chem. reactions of the active substances as reported by numerous investigators are summarized with an extensive bibliography. ALLEN S. SMITH.

The dielectric constant and the conductivity of ionized gases. TH. V. JONESCO AND C. MINET. *Compt. rend.* 191, 147-50 (1930). —The gas was ionized by a current of electrons in a glass tube provided at one end with a W filament, at 4 mm. from this with a grid, at 7 cm. a Cu ring and at 17 cm. an Al plate. When the ring and plate were at the same potential, ΔC and I were proportional to the electron current, ΔC being the change in capacity due to the change in dielec. const., and I the cond. of the gas. If λ is the wave length supplied by the oscillator and if the results are reduced to a current of 1 millamp., then I/λ^2 is neg. for wave lengths between 2 and 12 m. and then becomes pos. C/λ^2 is independent of the value of the accelerating potential. The wave length for which C/λ^2 changes sign is independent of the intensity of the current between 0 and 2 ma. The conductance, I/λ^2 is proportional to the intensity of the ionization. Its variation with λ is less simple. For an accelerating potential of 400 v. it has maxima at $\lambda = 2.4$ m., and 8 m., and a min. at 3.5 m. For greater wave lengths it decreases rapidly and approaches 0 for $\lambda = \infty$. For 100 v. I/λ^2 remains const. up to 6 m., and then approaches 0 rapidly as λ increases. R. H. LOWMEAD.

The hypothetical existence of Hertzian frequency resonators in water. P. GIRARD AND P. ABARIS. *Compt. rend.* 191, 1300-2 (1930). —The dielec. consts. of distilled water obtained by using wave lengths ranging from 4.15 to 60.50 cm. are recorded together with the dielectric consts. of 1.5×10^{-4} to 1.5×10^{-2} N₂O₄ at 10° for wave lengths ranging from 50.15 to 60.25 cm. The dispersion bands noticed by Weichmann and Frankenberg were not obtained in this region. The hypothesis for the existence of Hertzian resonators of a new type is found untenable because of the lack of confirmation of expl. data. T. W. LAIRD.

Excitation of fluorescence in benzene at -193° by monochromatic light. J. STARKIEWICZ. *Ann. phys. chim. phys. chim.* 1020, 27-41 of Pringsheim and Reimann. *C.* 4, 19, 420. Pringsheim. *C.* 4, 21, 1033. —The fluorescence spectra excited in C₆H₆ at -193° by the almost monochromatic radiation from electrodes of Zn ($\lambda = 2542$ and 2546 Å.) and of Cd ($\lambda = 2265$ and 2313 Å.) are in all respects similar to each other and to that obtained in similar circumstances by Reimann with a Hg lamp. This independence of fluorescence of wave length is due not as in the case of the vapor to collisions, but, as supposed by Pringsheim, to some intramol. mechanism. B. C. A.

The influence of the wave length of the exciting rays on the fluorescence spectrum of etioorphyrin. The structure of this spectrum in the infra-red and ultra-violet. J. ABARONIS AND CH. DUBRE. *Compt. rend.* 193, 14-15 (1931) of *Compt. rend.* 193, 321 (1931). With the wave length of the exciting rays varying from 355-546 mμ there were no noticeable variations in the no., position or intensity of bands in the fluorescence spectrum of etioorphyrin. T. H. RIVER.

Spectrophotometric study of the cupric-pyridine ion in aqueous solution. ALUMBAK AND TAMMIE. *Compt. rend.* 191, 93-100 (1930). —The absorption curves for various mists of Cu(NO₃)₂ and pyridine were detd. There is a max. in the curve, corresponding

to 20% $\text{Cu}(\text{NO}_3)_2$, showing the existence of $[\text{Cu}(\text{C}_4\text{H}_7\text{N})_4](\text{NO}_3)_2$. The stability const. for the deep blue complex ion is 3.1×10^{-6} at 18°. I. G. VANDER BROSCHIE.

The photoelastic dispersion of vitreous silica. I. N. G. FILON AND F. C. HARRIS. *Proc. Roy. Soc. (London)* A130, 410-31 (1931). The dispersion of the double refraction produced by stress has been accurately investigated with 2 blocks of vitreous silica. Results have been obtained for 23 standard λ lines between 4102 and 6105 Å. Values are given for the stress-optical coeff. at 6 different loads from 10 to 40 kg. The photoelastic dispersion curve shows anomalous dispersion. This may be due to natural periods of Si, as some of the irregularities in the curve correspond to lines in the Si spectrum. G. M. MURPHY.

Is the blackening and bleaching of exposed lithopone a photographic phenomenon. H. STORRE. *Z. Wiss. Phot. (Schaum-lehrschrift)* 20, 209-18 (1930).—An investigation is made of the light and dark reaction of ZnS and lithopone under different conditions. Phototropism is termed a photochem. reaction in which, by light exposure, a solid acquires a darker color which disappears again in the dark. S. compared the behavior of lithopone in light and dark with well-known photographic reactions and found differences which show that lithopone is not a photographic substance. A. P. H. T.

Behavior of dispersed systems in filtered ultra-violet light. KARL BORGMANN. *Süddeut. Apoth.-Ztg.* 71, 70-2 (1931).—A general discussion. W. O. I.

Luminous artificial filaments (U. S. pat. 1,791,199) 23. Geochemical distribution laws and cosmic frequencies of the elements (GOLDSCHMIDT) 2. Absorption spectra and the constitution of the desoxybenzoin series (SZCZÓ, OSTINELLI) 10. New photochemical effect on celloidin paper as a radiation problem (MENES) 5. Spectrography of the flavone series (HATTORI) 10. Absorption spectra of aroxy derivatives and analogous compounds (SZCZÓ, OSTINELLI) 10. Spectrochemical study of amino acid anhydrides (ASAHINA) 10. The influence of light on the methylene blue reduction (TAMIYA, *et al.*) 10.

BOHR, NIELS H. D.: *Atomteori og Naturbeskrivelse*. Copenhagen B. Lunos. 100 pp.

CHAFFUTS, JAMES, BERGET, ALPHONSE, AND LAMOTTE, MARCEL: *Leçons de physique générale*. T. IV. Ondes électriques, radioactivité, électro-optique. Paris: Gauthier-Villars et Cie. 214 pp.

Handbuch der Astrophysik. Edited by G. ENRIKHARD, A. KOHLSCÖTTER AND H. LUDENDORFF. Band III. Grundlagen der Astrophysik. Hälfte I. W. WESTPHAL: Wärmestrahlung. E. A. MILNE: Thermodynamics of the Stars. A. PANNEKOFK: Die Ionisation in den Atmosphären der Himmelskörper. S. ROSSELAND: The Principles of Quantum Theory. 473 pp. M. 74; bound, M. 77. Hälfte 2. W. GROTRIAN: Gesetzmässigkeiten in den Serienspektren. O. LAPORTE: Theorie der Multiplettspektren. K. WURM: Bandenspektren. E. A. MILNE: Theory of Pulsating Stars. 353 pp. M. 58.80; bound, M. 62. Berlin: J. Springer.

MARCH, ARTHUR: *Die Grundlagen der Quantenmechanik*. 2nd ed. of "Theorie d. Strahlung und d. Quanten." Leipzig: J. A. Barth. 293 pp. M. 24; linen, M. 26.

ROSSELAND, SVEIN: *Astrophysik auf atomtheoretischer Grundlage*. Band XI of "Struktur der Materie." Berlin: J. Springer. 252 pp. M. 19.80; bound, M. 21.20.

4—ELECTROCHEMISTRY

COLIN G. FINE

Recent developments in the design of electric annealing furnaces for metals. W. MECKENHAUSER. *Metal Ind. (London)* 38, 159-62 (1931). E. H.

1930 steel plant electrical developments. A. F. KENYON. *Blast Furnace Steel Plant* 19, 245-8, 254 (1931). E. H.

Development of chromium plating. J. G. ROBERTS. *Metal Ind. (London)* 38, 167-8, 194 (1931). E. H.

Industrial electric heating. J. I. BERNARD. *Elec. Eng. Australia and New Zealand* 7, 365-8 (1931).—A review. C. G. F.

The production of high-manganese slag in the electric furnace. T. L. JOSEPH. C. E. WOOD AND E. P. BARRETT. *Bur. Mines, Rept. of Investigations* 3080, 9 pp (1931).

—Flec. furnace tests made with 150 lb charges of spiegel, contg. about 15% Mn and 0.5% P, indicate that 2 types of high Mn slag may be produced from metal of this type. When SiO_2 was used as a thinning agent, the slag had an av. compn. of MnO 68.1.

FeO 5.2, SiO₂ 14.3, P 0.11 and Al₂O₃ 8.2%. With the exception of frothy slags, Al₂O₃ worked as well as SiO₂ as a thinning agent, the resulting slag being of the following composition: MnO 69.6, FeO 0.1, SiO₂ 8.7, P 0.16 and Al₂O₃ 13.0%. The SiO₂ and Al₂O₃ may be varied over a wide range provided their sum is about 22% of the slag. A continuous cycle gave better results than a batch operation. From 85 to 97% of the Mn could be recovered in the slag, 3-6% in the metal, and 6-8% was lost. More strongly reducing conditions can be established in the elec. furnace than in the open hearth, more positive control of P and Fe in the slag being thus permitted and the treatment being reduced to about 1/2.

ALDEN H. EMERY

The coreless induction furnace in a new role. A. G. ROBIETTE *Iron and Steel Ind. and British Foundryman* 4, 125-9 (1931), cf C A 24, 5638—Expts. were made to det. the possibilities of refining low grade steel scrap of various compos. and pig irons, the extent of the elimination of the elements S, Mn, P, C and S under varying conditions, and the speed of their elimination. Si is removed as rapidly as ore is charged. P elimination is rapid and complete under suitable slag conditions. C removal is rapid in both the slag reaction process and the air blast method. S removal is generally erratic, it increases with the CaF₂ addn. Products of deoxidation are removed in a few sec. through rapid circulatory motion of the charge. An auxiliary arc above the slag surface to keep it fluid is suggested.

C. H. LORIG

Hydrometallurgy of copper at the Bsgdad property. CHARLES T. BAROCH *Trans. Am. Electrochem. Soc.* 57, 205-29 (1930)—The ore minerals are principally chalcocite and chalcocite. Cu in the ore averages 1 to 1.25%. After roasting the ore it was leached with spent sulfate electrolyte, Cu extn. as high as 91% resulting. The soln. builds up in Cu to a concn. of 60 g./l. and more, the free H₂SO₄ dropping to 10 g./l. By diverting part of the spent electrolyte into this soln. coming from the extn. plant, the Cu is brought down to 50 g./l. and the acid up to 25 g./l., an electrolyte of good elec. cond. resulting. The c. d. of the Cu cells is 2.1 amp./sq. dm. Cell voltage has ranged from 1.97 to 2.12 v.

C. C. F.

The electrolytic zinc plant of the Evans-Walloway Company at East St. Louis, Ill. U. C. TAYNOR AND DANIEL BOSQUET *Trans. Am. Electrochem. Soc.* 57, 241-53 (1930)—The Joplin district will probably continue for many years to be the chief Zn producing district of the U. S. Hitherto, all Zn from this field has been produced by the old process of retort smelting. The paper describes the first plant to treat these ores electrolytically. Flotation concentrates are roasted in Wedge furnaces, the roaster gas being converted to H₂SO₄ by the contact process. After passing over magnetic separators, the cones are leached in spent electrolyte (29% H₂SO₄). The heat of the reaction raises the temp. to the b. p. Residues contg. Pb, Ag and Au are shipped to the Pb smelter. Leached soln. is purified with Zn dust to ppt. Cu, Cd, Co and Ni. The Cd is sep'd from the Cu and recovered electrolytically, while the Cu goes to a Cu smelter. In the Zn cells, Al cathodes are used at a c. d. of 10.7 am./sq. dm. Anodes are in grid form and are composed of Pb alloy contg. 1% Ag. During electrolysis, MnO₂ is ppt'd, a high grade product free from Fe, Cu, Ni, etc., being the result. The electrolytic Zn produced is 99.99% pure. The daily capacity of the plant is 50 tons of electrolytic Zn and 1/2 ton of Cd.

C. C. F.

Germanium in relation to electrolytic zinc production. U. C. TAYNOR AND E. T. CLAYTON *Trans. Am. Electrochem. Soc.* 57, 279-88 (1930) *Metal Ind. (London)* 37, 229-30, 326 (1930)—Certain impurities in the Zn soln., if not removed, will deposit with the Zn and lower the overvoltage below the critical point. In the 'high acid, high density' process Sb, Co, Ni and As, which are objectionable impurities, are readily removed. However, a hitherto unsuspected impurity which gave rise to low current efficiencies was, after careful investigation, found to be Ge. Concns. as low as 1 mg./l. are sufficient to cause serious losses in Zn cathode efficiency. Ge may be adsorbed and removed from soln. by means of Fe(OH)₃ ppt'd by ZnO in the leaching operation.

C. C. F.

The cathode precipitate as a function of exterior factors in solution of zinc sulfate. A. GLAZUNOV AND K. RÖSNER *Chem. Obozr.* 5, 221-3, 233-63 (263 English) (1930)—The theoretical conditions of Zn pptn. from ZnSO₄ soln. and the influence of various factors were investigated. High c. d., low temp., low H⁺ ion concn. and the presence of clay particles were favorable for the formation of a smooth and thin deposit. The concn. change from 25 to 110 g./l. Zn has practically no influence on the type of the deposit. The expts. are elucidated by numerous diagrams.

F. KUČKA

Some experiments on the electrodeposition of gold from alkaline cyanide solutions. E. T. DUNSTAN *J. Chem. Met. Mining Soc. S. Africa* 31, 118-27 (1930)—No evidence of regeneration of cyanide by electrolysis of a working cyanide soln. was found, in fact,

a further loss occurred. The use of a rotating cathode does not materially increase the current efficiency, which is extremely low, because of the dil. solns. employed. Electrolytic methods cannot economically improve upon Zn prim. It is doubtful if space and power would be saved, as the soln. required 4 passes through the cell. The rotating cathode has a depolarizing effect, probably because the friction of the soln. keeps the surface free from bubbles of gas, which enables a lower voltage and higher c. d. to be employed. Also in *Metal Ind.* (London) 33, 215-7 (1931) ALDEN H. MERT.

A new abnormality in the properties of aqueous solutions of cadmium salts. FLANKER TESCHÉ AND PIERRE J. VAN RYSSSELBERGHE. *Trans. Electrochem. Soc.* 59 (preprint) 5 pp (1931) —Electrolysis of aq. solns. of Cd²⁺ yields cathodic deposits containing a large amt. of adsorbed Cd₂. CdBr₂ behaves in the same way while CdCl₂ yields normal deposits of pure Cd. The phenomenon is ascribed to the adsorption of neutral molecules. An explanation for the difference of behavior of the solids and the bromide on one hand, and of the chloride on the other hand, is suggested. C. G. F.

Cadmium as a rust preventive. BERNHARD FLÄNKE. *Z. Elektrochem.* 37, 33-40 (1931) cf. C. A. 24, 330 —Cd is successfully deposited from an alk. cyan. bath at 40-60° and at current densities of 1-1.5 amp. per sq. m. with or without the addition of colloids. In acid baths conditions must be carefully controlled. With increasing temp. and c. d., the current efficiency increases, but the grain size of the deposit also increases. The best results are obtained at 40° with a c. d. of 100-150 amp. Acid baths are non-poisonous, and in them there is no decomposition of the anode. With increasing mol. wt. of the electrolyte the grain size of the Cd deposit falls. This is greatest in HClO₄ solns. and smallest in phenolsulfonic acid solns. Bright deposits can be obtained from fluosilicic and phenolsulfonic acid baths with added colloidal material, but such deposits are quite brittle and show poor adherence to the ground metal. H. S.

Addition agents in bismuth electrolytes. EDWARD F. KERN AND THOMAS R. JONES. *Trans. Am. Electrochem. Soc.* 57, 233-72 (1930) —Solns. of Bi³⁺ in benzoic, fluoroboric, cresol sulfonic, perchloric, ferrous, sulfite and chloride were tested for electrolytes of Bi, the result being the selection of BiCl₃ solns. contg. NaCl, CaCl₂ or MgCl₂ as the most suitable electrolytes. The addn. agents tried were gum arabic (sulfite residues), gum arabic, tannin, alone, alsoe curcuma, hydroquinone, benzoic acid, pyragallol and resorcinol. Pyragallol and resorcinol added to the electrolytes produced the brightest and densest deposits of Bi. hydroquinone was next in order, benzoic acid next, and the other addn. agents were not beneficial. C. G. F.

High-speed nickel plating as practiced in England. ERNEST R. CANNING. *Trans. Electrochem. Soc.* 59 (preprint) 4 pp (1931) —In England current densities of 2.5 amp. sq. dm. and over are used. Continuous filtration of the electrolyte, kept at 32° to 33°, is necessary to eliminate suspended foreign matter and particles. Because of considerable loss in throwing power, C. advises against very high current densities. The preferred ρ_a is 5.6 to 5.8. The cathode efficiency is 91 to 97%. A plate 0.023 mm. thick is obtained in 45 min. with a variation in thickness of not more than 45%. For adequate protection of the steel, a Ni plate of at least 0.023 mm. is necessary. For sub-tropical countries composite deposits of Ni-Cu-Ni are recommended. C. G. F.

Chromium plating. E. J. DOUSS. *J. Soc. Chem. Ind.* 49, 161-5T, *Metal Ind.* (London) 33, 43a-6, 46d-60 (1931) —A brief history of modern practice. The tendency in the U. S. is to aim solely at cost, and to overcome the disabilities attendant in high cond.-burning on the outstanding points of a recessed article, by the use of auxiliary cathodes or "robbars." In England research has been centered on increased "throwing power" without auxiliary cathodes or the use of specially shaped anodes. In England flat lead anodes are used. For articles subjected to atm. conditions a thick undercoat of Ni is essential. It is well to have the plating vats lined with hard Pb. This obviates any possibility of the continued formation of Cr(CrO₄)₃ and yet cannot cause any burning of the work from too large an anode surface, because of the screening action of the glass lining fitted to the latest types of vats to protect the Pb from mech. injury. The matters of temp., c. d., and prepn. of the articles for Cr plating and the stripping of faulty deposits are discussed. W. H. BOYER.

Variations in the electromotive force during the formation of alloys by the wet method. E. PACE. *Gazz. chim. ital.* 60, 811-8 (1930) —In connection with studies of Marrocchelli on the formation of binary metal alloys by the wet method, the phenomenon was examined from the electrochem. point of view. Similar expts. have been carried out by Tammann but under different conditions and with a different object (cf. C. A. 14, 671). The present paper deals with variations in the e. m. f. of the voltaic couple: A|A⁺ soln. |B, during the entire transformation in the following systems: Co-Sn, Co-Sb, Sb-Sn, Sn-Cd, Ag-Cd, Au-Sn and Ag-Sn. A high initial e. m. f.

would be expected because of the difference between the 2 potentials $A|A^+$ soln., and $B|A^+$ soln., where the first remains const. throughout and the 2nd is electrochemically indefinite as a result of the nominal absence of B^+ ions and theoretically is infinite. Practically, however, the 2nd has a definite value either because of soln. of traces of B (by the action of atm O or by double decomposition with A^+ cations), or, in case the B electrode is completely passive, because of the oxidizing potential which it acquires in contact with the soln. The app. is described in detail. In the $Cu-Sn$ system, during the formation of Cu_3Sn (cf *Ber* 27, 630(1907)) the $e.m.f.$ at 100° was initially 0.134 and thus rose slowly to a max. of 0.214. In the $Sb-Sn$ system the $e.m.f.$ at ordinary temp. with open circuit varied from 0.31 to 0.33, while with a closed circuit it diminished gradually to a run value of 0.80 after several hrs., and thus again increased to 0.300 when the circuit was opened. In the $Au-Cd$ system the $e.m.f.$ at ordinary temp. was 0.73 and at 100° was 0.77, and this diminished rapidly with the circuit closed until it reached 0.12 after 1 hr. at ordinary temp. and 0.013 after 20 hrs., while at 100° it diminished to 0.06 after only 15 min. In the $Ag-Cd$ system, the initial $e.m.f.$ at ordinary temp. was 0.67 and this increased slowly to a max. of 0.73. With the couple in a closed circuit at ordinary temp. the $e.m.f.$ diminished slowly from 0.72 to 0.42 after about 30 min. and reached 0.01 after 2 hrs. In the $Au-Sn$ system at ordinary temp., the initial $e.m.f.$ was 0.39 and diminished only to 0.37 after 6 hrs., while with the circuit closed it diminished from 0.37 to 0.05 after 20 min. When the temp. was increased to 100° , a closed circuit for only 3 min. was sufficient to reduce the $e.m.f.$ from 0.36 to 0.06. In the $Ag-Sn$ system at ordinary temp., the $e.m.f.$ was approx. 0.35, and with the couple in a closed circuit this diminished to 0.01 after only 2 min. At 100° the behavior was substantially the same, except that the changes were accelerated. C. C. DAVIS

Electrodeposition of cobalt-nickel alloys. II. S. GLASSTONE AND J. C. SPEAKMAN. *Trans. Faraday Soc.* 27, 24-35(1931).—The investigations of the compn. of alloys deposited from mixed buffered solns. of Co and Ni sulfates under various conditions reported previously (*C. A.* 25, 459) have been extended to higher temps. approx. 50° and 90° . In the present work it has been found that the variation with increasing $e.d.$ of the compn. of the $Co-Ni$ alloys is very different at high and at low temps. At 15° the compn. of the alloy is almost independent of $e.d.$ until the rate of removal of ions from soln. by electrodeposition exceeds the rate of diffusion of these ions to the electrode, at 50° the proportion of Co increases rapidly with increasing $e.d.$ before a const. compn. is reached (cf *C. A.* 21, 3025, 22, 3555). The exptl. conditions in the present work were as reported in Part I, and graphs of the atomic % of Co plotted against $e.d.$ in amps. sq. dm. are given. The relative tendency for Co and Ni to deposit is independent of the H ion concn. of the electrolyte. The alloys deposited at very low $e.d.$ s contain decreasing proportions of Co as the temp. is raised; this proportion is, in general, less than that in the soln. for depositions made at 90° . With increasing $e.d.$ the fraction of Co in the alloy increases rapidly until a max. Co content, independent of $e.d.$, is attained; this max. decreases with increased temp. The influence of Co on the deposition of Ni at 50° and 90° is very similar to that of Fe at 15° and 50° , resp. In general, the current efficiency increases, for the reason that the deposition potential decreases more rapidly than the H overvoltage, with increasing temp. The theoretical basis of the results is discussed in terms of the deposition potentials and overvoltages of the metals concerned. EDWARD B. SANIGAR

Polarographic studies with the dropping-mercury cathode. IV. Positive and negative maxima on current-voltage curves. J. HEVROVSKÝ AND M. DILLINGER. *Collection Czechoslov. Chem. Comm.* 2, 626-34(1930).—The anomalies observable on electrocapillary curves obtained by the drop wt. method have been shown (*C. A.* 24, 2670) to be indicated on the current voltage curves by prominent max. which are given under conditions of imperfect polarization of the dropping Hg cathode. Theoretical treatment of these anomalies is given in the present paper, and it is shown that the presence in soln. of a highly adsorbable substance, e.g. an org. dye, prevents this imperfect polarization and therefore, the anomalous shape of the current voltage curve. The max. occurring on current voltage curves when solns. contg. electroreducible substances are electrolyzed with the dropping Hg cathode are termed "positive" or "negative" according as they are given at potentials more pos. or more neg. than the abs. electrocapillary zero ($s.e.$, -0.56 v. from the N calomel zero). The shape of the electrocapillary curve, simultaneously derived from the polarized dropping- Hg cathode, indicates the sign of the max. Just as the pos. branch of the electrocapillary curve is affected by the presence of adsorbable anions and the neg. branch by adsorbable cations, the presence of such anions has been shown to affect the pos. max., while neg. max. have been shown to be sensitive to cations. The current voltage

charge. The formation of the oxide, Ag_2O , on a Ag electrode has thus demonstrated, and similar expts. with accumulator plates have yielded results supporting the theory that Pb suboxide is formed on the neg. plate at discharge, but no definite evidence of the presence of Pb_2O_3 on the charged pos. plate could be obtained (cf Féry, *C. A.* 13, 1576; Féry and Chéneveau, *C. A.* 20, 1181). In the Edison accumulator, the active material on the charged Ni plate is mainly NiO , if the alkali concn. exceeds about 6 N .
W. H. BOYCE

The Volta effect in electrochemistry. OSCAR SCARPA. *Atti III congresso naz. chim. pura applicata* 1930, 83-94.—S. has devised an electrometer in which the displacement of a metallized quartz fiber suspended between two metal plates measures the p. d. between the plates. With this instrument he has measured the p. d. between a Zn and a Cu plate when they have been attached to the Zn and Cu electrodes, respectively, of a Daniell cell. He has also measured the p. d. between a Zn and a Cu plate connected by a Cu wire and between two Cu plates connected to the electrodes of a Daniell cell. In the first case the e. m. f. was 0.21 v., in the second 0.85 v., and in the third 1.08 v. The first case S. calls the true or surface Volta effect and ascribes it to the e. m. f. of contact between the electrodes and the solns. and between the two solns. The second or internal Volta effect is the contact potential of the two metals. J. B. ACSTIN

Formation of methane during the electrolysis of potassium acetate and the mechanism of Kolbe's electrolysis. S. N. SHUKLA AND O. J. WALKER. *Trans. Faraday Soc.* 27, 35-40 (1931).—Evidence has been obtained that CH_4 is formed at the anode during the electrolysis of aq. acetate solns. under certain conditions. It is suggested that the presence of CH_4 has been overlooked by previous investigators, because their analyses of the gaseous products of electrolysis have been confined almost entirely to the mixed anode and cathode gases (i. e., to mixts. contg. a large proportion of H_2) resulting from electrolysis at high c. d. If it is assumed that only C_2H_6 and H_2 are present in the combustible residue, after the removal of the other constituents the presence of small amounts of CH_4 may remain undetected since $2\text{C}_2\text{H}_6 = \text{C}_2\text{H}_6 + \text{H}_2$; i. e., the CH_4 simply acts as a mixt. of C_2H_6 and H_2 . In the present work the anode gases were analyzed separately from the H_2 given at the cathode. The combustible residue of the anode gas after the removal of the O_2 and CO_2 in a Bone and Wheeler app. over H_2SO_4 was transferred to an evacuated tube and cooled in liquid air to remove the C_2H_6 , the contraction after cooling being a measure of the C_2H_6 present. After some time the uncondensed gas was pumped off by a Töpler pump and exploded. The ratio, contraction on explosion (C)/amount of absorption by potash (A) was approx. 2 whereas the C/A ratio for the condensed gas was 1.25 (the theoretical ratio for C_2H_6), showing that CH_4 was present in the above anode gases. In the electrolysis a Pt wire anode (4 cm. long, 0.1 cm. diam.) was used, and the anode gases were collected separately. Solns. of K acetate (contg. an equiv. amount of acetic acid) of different concns. were electrolyzed at various c. d.s. at ordinary temp. A table of exptl. results is given. The results show that under certain conditions CH_4 is present with C_2H_6 in the anode gases produced during the electrolysis of K acetate solns., the conditions favorable to the formation of CH_4 being low c. d. and high concn. of electrolyte. The quantity of CH_4 was appreciable only when the c. d. was less than about 10 ma./sq. dm. In the most favorable instance, with a 3.5 N K acetate soln. plus 3.5 N acetic acid, at a current of 1.5-2.0 ma., almost twice as much CH_4 as C_2H_6 was produced, the total exptl. results for this soln. being CO_2 , 67.3; O_2 , 0.4; CH_4 , 19.8; C_2H_6 , 19.5 and residue 2.6%. The results also show that CH_4 is formed only along with C_2H_6 , no CH_4 is evolved so long as O_2 is the main gaseous product. With ordinary c. d.s. such as are used in the electrolytic prepn. of C_2H_6 , no appreciable amt. of CH_4 is produced. The mechanism of Kolbe's electrolysis and the recent work of Fichter are discussed. The abnormally high value of 2.54 v. for the discharge potential of the acetate ion (as put forward by Preuner and Ludlani, *C. A.* 2, 12) is not correct, C_2H_6 is evolved below this potential.

EDWARD B. SANIGAR

Self discharge of dry cells. SABURO MAXIMO. *J. Soc. Chem. Ind., Japan* 33, Suppl. binding 541 (1930).—To investigate the mechanism of self discharge a positive electrode only was immersed in an electrolyte, and the same electrode was immersed as an anode with the usual neg. electrode. Decrease of e. m. f. and capacity during open circuit occurs first at the anode, then gradually at the cathode, the degree depending upon the amt. of Fe contained as impurity in the MnO_2 used. E. M. S.

High-speed photographs of arcs in liquids. CHESTER LICHTENBERG. *Elec. World* 97, 321, 323 (1931).—Arcs sprung in insulating liquids ordinarily persist 0.01-0.1 sec. A high-speed, successive-image camera for studying these arcs is illustrated. It clearly records 24 individual successive images at a rate of 550 images per sec. The arcs were

started by opening a single phase, 6000-v., 25-cycle a. c. carrying 50 amp at 0.45 power factor. The same general phenomenon is observed in all the photographs. The arc starts at the contacts, and immediately oil or other liquids in the vicinity of the arc is gasified. A gas globule forms around the contacts. As the contacts continue to sep., the gas globule expands as a sphere till the sides of the containing vessel or a baffle plate alters its shape. Simultaneously the oil level is raised a distance proportionate to the size of the gas globule. Finally the arc punctures the globule, releasing the pressure within. Sometimes the release is violent. Then the oil level is distorted and broken, some oil being thrown upward. The remainder descends, cools the arc and extinguishes it. W. H. BOYNTON

A study of the physical processes in the so-called electrical purification of gases. 11. The action of the electrical wind (Cottrell process). R. LADENBURG AND W. TIEZIE. *Ann. Physik* [5], 6, 581-621 (1930), cf. *C. A.* 24, 3147. —This is an investigation of the effect of the elec. wind (EW) on the velocity of dust particles being pptd. from a gas by elec. pptn. Using the corona discharge from a neg. charged wire to a grounded wall, the authors have found that for particles of diam. $\leq 10^{-4}$ cm. the EW is the detg. factor on the velocity (v) which at 10^{-4} amps. and 5×10^{-4} amps. glow-discharge point is 0.56 and 1.23 m. sec., resp. v decreases with distance from glow-discharge point, being for 5×10^{-4} amps. at 3 mm. 1.20 m./sec. and at 60 mm. 0.80 m./sec. For particles of diam. = 10^{-4} – 10^{-3} cm. the EW is still an important factor. The measurement of v made by comparing the effect of a current of air of known velocity with that of the EW on a stream of CO_2 , tobacco smoke or P_2O_5 in a specially constructed cylindrical pptn. chamber was recorded photographically. Description and drawings of the app. and photographs of the effects are given. I. R. SCHERER

High-voltage surge testing—the cathode-ray oscillograph. F. D. FIELDER. *Elec. J.* 28, 106-7, 119 (1931). —The cathode-ray oscillograph accurately measures voltage and time relationships. It depends for its operation on the influence of elec. fields upon an electron beam. The Dufour and Norinder types are illustrated and their operations outlined. The deflection of the cathode ray by an elec. field is a pos. motion, and oscillograms indicate it with an accuracy of 2%. W. H. BOYNTON

1930 developments in electrical equipment for steel mills. H. A. WATNE. *Iron Steel Eng.* 8, 21-5 (1931). —The synchronous motor has very largely superseded the induction motor for const. speed mill drives, except where flywheels are necessary. The year's developments on d. c. main drives, switchgear, motor room ventilation, auxiliaries, vacuum tubes and elec. furnaces are noted. W. H. BOYNTON

Low-temperature carbonization and the production of electricity in Germany (ROSEN) 21. Pb (THOMSON) 9. The calculation of chemical, electrical and gas mechanical matter-, energy- and form transformation by space energetics (DALLWITZ-WEGENER) 2. Some properties of protective films on metals (HEDGES) 9. Refractory materials for electric furnaces (SEARLE) 19.

Handbuch der technischen Elektrochemie. Band I, Hälfte 1. Die technische Elektrolyse wässriger Lösungen. Edited by VICTOR ENGELHART. Leipzig. Akad. Verlag. 613 pp. M. 56, bound, M. 58. By subscription, M. 50.40, bound, M. 52.20.

Dry cell electric battery assembly. SIEMENS & HALSKE A-G. Brit. 335,758, Oct. 19, 1928. Structural features.

Electric dry cell battery assembly. JOHN S. ZOOK (to Burgess Battery Co.). U. S. 1,790,784, Feb. 3. Structural features.

Storage battery. SOLOMON L. VAN MEYER, JR. U. S. 1,791,152, Feb. 3. Structural features.

Alkaline storage battery. J. J. DRUMM and CELIA, LTD. Brit. 335,587, March 27, 1929. Ag or Ag oxide mixed with Ce oxide is employed as active material for the pos. electrode. Various details of prepn. are given.

Storage-battery plate. ALBERT S. HUBBARD (to Gould Storage Battery Corp.). U. S. 1,791,208, Feb. 3. In making a tubular container to form a part of a storage-battery plate and to contain a core surrounded by active material, a sheet of non-porous insulating material such as hard rubber stock is rolled out, laminae of porous material such as cloth, wood or felt are embedded in the sheet, and the latter is cut into a plurality of washer like sections which are then stacked and vulcanized.

Selenium cell. HANS THIRRING (to "Selenophon" Licht- und Tonbildges. m. b. H.). U. S. 1,790,850, Feb. 3. Various details of manuf. are described.

Accumulators. MAURICE MATROT Fr 692,917, June 27, 1929 The electrolyte of an accumulator having a Zn neg. electrode is composed e. g., of H_2SO_4 (45°Bé) 590 cc, ZnSO_4 3%, dimethylglyoxime 4, Na_2SO_4 50 and MgSO_4 20 g

Tin plating. SIMMONS & HALLSKE A-G Fr 692,813, Mar 27, 1930 In the electrolytic formation of coatings of Sn of any thickness alk. electrolytes and sol or insol. anodes being used, the content of free alkali is kept between the limits of 1.5 and 8%, so that no spongy formation is produced during the time necessary to obtain the desired thickness of Sn

Electrolytic production of chemical compounds. ANTONIO L. CLARIANA Fr 693,150, April 2, 1930 In electrolytic processes for the production of chem. compds. the cathode is made of or coated with a metallic sponge of finely divided metallic particles which are easily oxidized in contact with the air. The operation is intermittent, the cathode, when reduced by the H_2 produced, being brought into the air until again oxidized

Electrolytic apparatus for cleaning the surface of metals. THE BULLARD CO Fr 693,199, April 2, 1930

Electrolytic gas generator. PAUL HAUSMEISTER Ger 514,391, April 20, 1924

Electrolytic refining of copper. UNITED STATES METALS REFINING CO Fr 693,678, April 10, 1930 The cathode used in the refining of Cu have parallel corrugations extending in 2 different directions

Electrolyte for the production of iron or iron alloys. ORLANDO ORLANDI Fr 693,812, April 12, 1930 See Belg 360,116 (C A 24, 793)

Electrolytic manufacture of tungsten, etc. HELLMUTH HARTMANN Ger 514,365, Sept 11, 1928 W is manuf. by electrolyzing a soln. of WO_3 in fused alkali phosphate, with or without H_3PO_4 or other phosphates. An example is given. Other heavy metals, e. g., Mo, Ta, Vanil Nb, may be extd. similarly

Electrolytic zinc bath. GRISOPPE BIANCO U S 1,701,082, Feb 3 A bath suitable for use in coating iron or steel with Zn is formed from a Zn amalgam, alkali metal phosphate, H_2SO_4 , $\text{Al}(\text{OH})_3$ and water

Light metals. SOC ANON POUR L'IND CHIM A BALE Fr 692,491, Mar 21, 1930 In electrolyzers for the prepn. of light metals by electrolysis of their molten halogen salts a cathodic arrangement for gathering the light metal formed is used. It is composed of an exteriorly insulated metal piece in the form of a longitudinal channel or provided on its lower face with longitudinal channels, in elec. connection with the cathode which is placed below the said piece parallel to the direction of the channels and descends relatively low into the electrolyte

Magnesium. THE MAGNESIUM PRODUCTION CO., LTD (Pierre Camet Casse, inventor) Fr 693,387, Apr 4, 1930 In the manuf. of Mg by the electrolysis of a soln. of MgCl_2 and KCl , a part of the electrolyte is removed from the vat, and MgCl_2 is added so that the compn. of the electrolyte is maintained practically constant and the KCl follows a closed cycle

Purifying liquids. JEAN BILLET Fr 693,799 April 12, 1930 Liquids are purified electrolytically, while suppressing electroosmosis as much as possible, by using mean differences of tension of less than 20 v. (preferably 8-16)

Hydrogen from water. ALBIN WERLEN THIEBAUT KUENEMANN, ÉDOUARD SILBERMANN and CHARLES GOETSCH Fr 693,637, Jan 21, 1930 In the production of H_2 from water an anode composed of a Hg-Cu-Zn catalyst, a depolarizer composed of an alkali phosphate and an At cathode are used

Butyl alcohol. R. RILEY S. W. ROWELL and IMPERIAL CHEMICAL INDUSTRIES, LTD Brit 335,683 Aug 13 1929 BuOH obtained by the catalytic hydrogenation of aldol crotonaldehyde or butylaldehyde is purified by electrolytic reduction in aq. electrolyte either acid or alk. such as H_2SO_4 or NaHSO_4 soln. The product may sep. into 2 layers the upper of which when neutralized and distd. forms a distillate also sep. into 2 layers the upper of which contains the purified BuOH substantially free from aldehyde

Electric furnace for aluminum manufacture. PEDER ELIASSEN FRÖLAND Ger 515,220, Mar 6 1928

Induction furnace. F. MIL F. RUSS Ger 514,357, Dec 12, 1929

Electric induction crucible furnace. C. LORENZ A-G Brit 335,766, Oct 25 1928 Structural and elec. features

Electric-arc furnace suitable for making tungsten carbides. OSCAR L. MILLS (to Mills Alloys Inc.) U S reissue 17,954, Feb 3 Reissue of original pat No 1,719,553 (C A 23, 3863)

Heating elements for electric furnaces. F. MIL F. RUSS Fr 693,413, April 5, 1930

Apparatus for electrical precipitation of suspended particles from gases. HARRY WINTERMUTE (to Research Corp.) U S 1,791,008, Feb 3. Structural features.

Apparatus for electrical precipitation of suspended particles from gases. HARRY V. WELCH (to International Precipitation Co.) U S 1,790,961, Feb 3. Structural features.

Armoured electric deep-sea cables. FELTEN & GUILLEAUME CARLSWIDE A-G. Brit 325,504 Jan 23 1929. Various structural details are described of cables with a pressure protecting covering formed of profile wires and covered by a water tight sheathing of gutta-percha, gutta-guttsch or washed rubber.

Electric incandescent lamps. N. V. PHILIPS' GLOEI-LAMPEN-FABRIEKEN. Brit. 325,617, June 29 1929. Various structural details are described of a lamp which is adapted to be mounted within a parabolic reflector and the bulb of which is preferably colored by surrounding it by a template or stencil and spraying with a "Bakelite" lacquer. A yellow layer may be obtained by burning in Ag oxide.

Metal filament for incandescent electric lamps. C. SEVEN. Belg 370,823, July 31, 1929. The filament is formed of an alloy of W, Th, Ir, Ni and Cr, or of an alloy of W, Th and Zr. It has an elongated cross section, either an ellipse or a double parabola.

Lamp filaments. Soc. d'Etudes Scientifiques et Industrielles. Fr 693,275 June 1, 1929. Filaments for wireless lamps are made of Ba_3N_2 or Ba alloys such as Na Ba, Bi Ba or Sn Ba. Cf C A 25, 1170.

5—PHOTOGRAPHY

E. F. WIGHTMAN

Preparation of fine grain emulsions. E. FUCHS. *Phot Ind.* 28, 1172-4(1930).—The graininess observed in photographic materials is due to the clumping or aggregation of grains rather than to the individual grains themselves. Other things being equal, graininess is proportional to the size of the individual grains, but the prepn. of fine grain emulsions is largely a matter of colloidal principles aimed at preventing formation of aggregates of grains. In the ammonia process, substitution in part by Na_2CO_3 is recommended for the first ripening, followed by a quick second ripening of the well washed material. C. E. MEULENDYKE

Preparation of silver bromide emulsion. MARX REO, T. NISHIMURA, II HAYASHI-HARA, T. ASAI AND S. MURATA. *Repts. Imp. Ind. Research Inst. Osaka, Japan* 11, No. 13(1930).—The methods of prepn. of AgBr emulsion together with its properties are given. F. I. NAKAMURA

Solubility of silver bromide in photographic emulsion. MICHIO MIYATA. *Repts. Imp. Ind. Research Inst. Osaka, Japan* 11, No. 12(1930).—The use of the pptn. method in detg. the soly. of AgBr in photographic emulsions is given. F. I. NAKAMURA

Toning by addition. S. JASIEŃSKI. *Atelier Phot* 37, 113-5(1930).—Processes for toning Ag images fall into two groups: the substitution methods, in which the Ag of the image is changed to another form, as Ag₂S, and the addn. methods, in which some substance is added to the Ag, as the U toning process. The final deposit obtained when a HgCl₂ bleach is used for sulfide toning is partly sol. in Farmer's reducer. It cannot, therefore, be Ag₂S but is probably a complex Hg S compd. II PARKER

Alkaline selenosulfates and their use for the toning of silver images. A. SEYEWITZ. *Sci. ind. phot* [2], 1, 436, *Rev. franç. phot* 11, 332, *Photographe* 17, 475-6(1930).—Seleno-sulfates can be obtained from NH_4 , Li, Na and K sulfites. Because of its great soly., the last is the most interesting. On dilg. a concd soln. of K selenosulfate with water, Se is pptd., but on dilg. with a K_2SO_4 soln., the soln. remains colorless. Silver selenosulfate decomposes, readily pptg. Ag₂Se, hence, the soln. cannot be used to tone prints. Several prints may be fixed and toned simultaneously with a fresh soln. of hypo contg. in addn. K selenosulfate. After fixing 2 or 3 prints, however, clear whites are no longer obtained because of the decompn. of the Ag selenosulfate. L. A. STAIB, JR.

History of the invention of silver chloride developing-out paper of the Velox type. R. NAMIAS. *Il. prog. fot* 37, 397(1930).—In a discussion of this subject, Eder claimed that AgCl papers were originated by himself and Pizzighelli. Backeland replied that in his papers the emulsion was not washed, and that the present type of developing out paper therefore originated with the introduction of Velox. Backeland's claims are supported by N. C. E. K. MEES

Determination of so-called Ur-silver in photographic gelatin. H. H. SCHMIDT AND F. PRETSCHNER. *Z. wiss. Phot.* 28, 293-7(1930); cf C A 25, 42.—The previous

differentiation of the Ag content between atomic Ag and an organic Ag compd does not hold. The Ag is not the result of the action of AgNO_3 on gelatin, but is due to differences in the diffusion velocities of Ag and halogen ions within the gelatin during washing. Ag halide and Ag formed by exposure cannot be sepd with neutral $\text{Na}_2\text{S}_2\text{O}_3$. The best available substance is Na_2SO_3 . Its use, however, is limited to the investigation of AgCl and AgBr . A. P. H. TRIVELLI

Origin and the chemical nature of the silver content. H. H. SCHMIDT AND F. PRETSCHNER. *Z. wiss. Phot.* 28, 302-11 (1930).—The Ag content arises by hydrolysis through differences in the diffusion velocities of halogen and Ag ions and consists of Ag_2O . This Ag_2O has a tanning action on gelatin, forming a difficultly sol complex. The total amt. of detectable Ag_2O in the gelatin after washing depends upon the relations between the soly of Ag halide and the Ag₂O gelatin complex. With the same soly of the Ag complex, the Ag content increases with increasing soly of the Ag halide. The total amt. of Ag_2O formed very quickly approaches a limit when as much Ag_2O dissolves as is formed by hydrolysis of Ag halide. The Ag content resulting from AgCl is perfectly sol. AgBr and AgI were also studied. It is possible that the Ag_2O formed by hydrolysis influences ripening. In any case, Ag_2O is more easily reducible than Ag halide. A. P. H. TRIVELLI

Silver content and the fixing process. H. H. SCHMIDT AND F. PRETSCHNER. *Z. wiss. Phot.* 28, 297-302 (1930), cf. C. A. 24, 2009.—The soly of Ag is less in Na_2SO_3 than in $\text{Na}_2\text{S}_2\text{O}_3$ soln. Therefore, the total removal of Ag from a washed emulsion by Na_2SO_3 soln is due to a sol. Ag compd which has been shown to give Ag ions. The very small amts. of Ag left in the emulsion are of the same order as those obtained if pure gelatin is treated with thiosulfate and the sulfite complex compd solns. Very probably this Ag is Ag_2S . The soly of Ag in thiosulfate solns. is too great for the use of the latter in accurate analytical investigation. A. P. H. TRIVELLI

Mechanism of the color change of silver images due to selenium. A. SEIBWITZ. *Sci. and Phot.* [2], 1, 435 (1930).—Se dissolved in alkaline sulfites or sulfides can be used to tone Ag images. Sodium selenosulfate and sodium selenosulfide react with Ag to give Ag_2Se . This reaction has been verified by removing the image from plates and paper and detg. analytically the ratio of Ag to Se. A slightly larger percentage of Se was found in the selenosulfate toned image. Satisfactory agreement with the theoretical ratio was obtained for the selenosulfide image. L. A. STAIN JR.

Orthochromatism, panchromatism and dye filters. GULLIMOT AND BOESFELUG. *Photo. Revue* 42, 309-11, 325-6 (1930).—An estn. of the chromatic sensitivity of negative materials may be obtained: (1) by photographing a color chart with and without filters, (2) by means of the Fiedler Hecht sensitometer wedge (3) by means of spectrographic curves. 1 samples of each method are shown. H. D. RUSSELL

Glycocoll-copper for daylight filter. R. LUTHER. *Kinotechnik* 12, 453 (1930).—The advantages are: (1) it is easy to prepare in c. r. form and easy to compound, (2) it contains no volatile and easily oxidized material (3) it is neutral and does not attack the glass walls of the cell (4) it is practically stable with time, (5) the absorption is more independent of temp. The disadvantages are: (1) the soln. can develop mold, (2) it is sensitive to alkali and acid its prepn. being thus made more troublesome, (3) its soly is low at room temp. If necessary, a greater thickness of cell may be chosen. The extinction curve is shifted 4 or 10 m μ farther toward the red, and the density is less in the ultra violet than is that of the Davis Gibson filter. It appears to give a closer match to the curve of the ideal filter than does a CuSO_4 pyridine mannitol cell. O. E. M.

Raising the sensitivity of bichromated collod film. K. JACOBSON AND E. WAGNER. *Atelier Phot.* 37, 98 100 (1930).—The addn. of CeCl_4 to the sensitizing bath as advised by F. J. Tritton was not found to increase the sensitivity of bichromated gelatin. Bathing in a soln. of methylene blue after the usual sensitizing with $\text{K}_2\text{Cr}_2\text{O}_7$ gave increased red sensitivity. The loss in the green and yellow balances the gain in the red, so there is no gain in the total sensitivity. H. PARKER

Oleobrom. J. D. JOHNSTON. *Camera* (Dublin) 10, 235-6 (1930).—Technic of bromol making is simplified by the introduction of a new bromide paper. The developed print is bleached and dried in the usual way. The dry print is charged with ink from a special roller. The print is then placed in water and a clean roller run over the print. This rolling removes the ink from the highlights and deposits it on the shadows. After the proper contrast has been obtained, the print is removed from the water and is either dried or the tones altered by hand manipulation. R. A. PURDY

New photochemical effect on celloidin paper as a radiation problem. H. MEWES. *Z. wiss. Phot.* 28, 311-24 (1930).— AgCl papers exposed to different substances (ele-

ments of the periodic system) in a nonconducting box of ebony, or metal free, colorless paper maché, and developed in dil. rodinal developer (1:25) at 18° for 15 min., give pictures which show the effect of radiation and not the effect of emanation of H_2O_2 . This radiation M. calls *L* rays. Different properties of these *L* rays are discussed. A. P. H. TRIVILLI.

Testing wash water for hypo. *Atelier Phot* 37, 117 8(1930).—Several methods are described for testing for the presence of $Na_2S_2O_3$. (1) A dil. (light red) soln. of $KMnO_4$ is decolorized in less than a min. after the addn. of water contg. $Na_2S_2O_3$. (2) Two cc. of a 1% soln. of $AgNO_3$ acidified with HAc , mixed with a glass full of the wash water, give a brown color if $Na_2S_2O_3$ is present. (3) A drop of satd. $HgCl_2$ soln. added to a small amt. of the wash water gives a brown color with $Na_2S_2O_3$ present. The method of Crabtree and Ross using a soln. of $HgCl_2$ and KBr is also described. H. PARKER.

Ferrous oxalate developer. *Séanl. Atelier Phot* 37, 100 2(1930).—For overexposures, the FeC_2O_4 developer gives better results than ddd. or restrained org. developers, and the time of development is normal. Two stock solns. are used: (f) 1:3 (33%) $K_2C_2O_4$ and (B) 1:1 (50%) $FeSO_4$ with a drop of H_2SO_4 . Three parts of A to one part of B gives the strongest developer; increasing the proportion of A weakens it, or KBr may be added to correct overexposures still further. One or 2 drops of hypo increases the energy, but an excess causes fog. A good reducer may be made from the used developer. Fifteen g. of oxalic acid is dissolved in a liter of the heated developer. The green crystals of ferric potassium oxalate formed when the developer cools are made up as a 1:20 soln. and hypo is added equal to 5 times the wt. of the green crystals whose action is between that of $K_2C_2O_4$ and persulfate. H. PARKER.

Is the blackening and blanching of exposed lithopone a photographic phenomenon? (SROON) J. The Spierer lens and what it reveals in cellulose and protoplasm (SEIFRIZ) H.A.

Photographic coating sensitive to electronic discharge. AUSTIN G. COOLBY, U. S. 1,701,051, Feb. 3. Finely divided conducting particles such as metal are distributed through a sensitive coating to render it suitable for use in recording app.

Three-color cinematograph film. LEON J. B. DIERCK, U. S. 1,700,978, Feb. 3. A blue image is printed on one face of a film coated on one side with a gelatin bromide layer and on the other with a neutral gelatin layer; the film is coated with a soln. of a dichromate, and there are printed simultaneously on the sides of the film a yellow and a red image, resp., the film being then dried and provided with a colored gelatin strip on each side.

Production of colored photographic and cinematographic positives. L. J. DASSONVILLE, Belg. 372,003, Aug. 31, 1930. A Ag salt image is first produced in the sensitized emulsion. After resensitization by means of a salt of another metal, a 2nd image is produced. It is obtained colored by pptn. of a colored salt.

Photographic plates and films. I. G. FARBENFABRIK A.-G., Fr. 693,659, April 10, 1930. Anti-halo layers for photographic plates and films contain fuchsonic dyes with one or more COOH groups in the mol. Examples are given of the use of aurindicarboxylic acid, aurinmonocarboxylic acid and hydroxycarboxyfuchsonic.

Light-sensitive layers. KALLER & CO. A.-G., Fr. 693,032, Mar. 29, 1930. Light-sensitive layers are prepd. on appropriate supports using diazo compds. of aromatic amino-acylamines of the general formula NH_2 , $R-NH-CO-X$ and NH_2 , $R-NH-CO-NH-R'$, NH_2 , in which R and R' are substituted or unsubstituted groups and X is an aliphatic, aromatic or aliphatic aromatic group and the CO may be replaced by CS . Examples are given.

Protecting surfaces. VLADIMIR REVOŠ, Fr. 692,800, Mar. 26, 1930. Fragile surfaces such as photographic negatives or papers, films, pictures, etc., are protected by a thin layer of a transparent soln. of celluloid, etc., in an appropriate solvent such as $AcOAm$.

Printing molds. ERNST BECHERT, Fr. 692,794, Mar. 26, 1930. In the production of printing molds, safranine or like substance is added to the light sensitive layer to neutralize the lateral or undesired illumination.

6—INORGANIC CHEMISTRY

A. E. MIDDLETON

Researches on the role of water in the salts' aquo combinations of ruthenium-IV-ammine. R. CHIRONVAT. *Compt rend* 191, 1453-5 (1931) —In partial support of his theory of hydrolysis Werner made use of certain transformations of salts of Ru-IV nitrosylhydroxyl-triammine which are caused by suitably heating the salts with acid, namely: $X_3[Ru(NO)(OH)(NH_3)_3] (I) \rightarrow X_3[Ru(NO)(H_2O)(NH_3)_3] (II) \rightarrow X_3[Ru(NO)(C_2H_5)_3] (III)$. In his investigation of the role of water in salts of type II, C. prepd 2 new salts having the groups $Pu(NO)_2 \cdot n_1 Cl_2 \cdot H_2O$ and $Ru(NO)pyr \cdot Cl_2 \cdot H_2O$ (n_1 = ethylenediamine and py = pyridine). The properties of these complex aquo salts indicate that they are really hydrochlorides of hydroxides having the following structures: $Cl_2[Pu(NO)(C_2H_5)_3](OH)(H_2O)]$, $[Ru(NO)_2 \cdot n_1(C_2H_5)_3](OH)(H_2O)]$, and $[Ru(NO)pyr_2(C_2H_5)_3](OH)(H_2O)]$. The mol of HCl is joined to the O of the group $(HCl \rightarrow O-A)$ by a secondary valence, the O possessing tri-coordination.

R. H. LOWHARD

A new method for the conversion of alkali sulfate to alkali chloride for use in sulfate analysis. WERNER MYRUS. *Sprechsaal* 63, 972-5 (1930) —A new method for converting the alkali sulfate to the chloride in the quant. detn. of K and Na in sulfates to replace the more bothersome $BaCl_2$ method consists in a double evap. of the finely pulverized neutral sulfate with $N_2H_4 \cdot 2HCl$.

C. H. LING

The reaction between selenic acid (alkali iodates) and hypophosphorous acid, phosphorous acid or their salts. (III). V. HOFERKA. *Collection Czechoslov Chem Comm* 2, 669-25 (1931) (in French), cf. C. A. 25, 458 —This part deals with the oxidation of the lower acids of P by HIO_3 in the presence of substances capable of removing, by the formation of insol. iodides, the I produced by the reaction between the HIO_3 and the HI formed during the oxidation of the phosphorous acids. Salts of Pb, Ag, Hg^{++} and Cd were used, but with Pb and Cd salts free I was always formed. Excess of the perchlorates of Ag and Hg^{++} was finally used. The work was carried out with acidified (H_2SO_4) alkali phosphite or hypophosphite salts to which KI and excess metallic perchlorate were added. With Ag and Hg^{++} perchlorates the corresponding iodides were formed, so that very little direct oxidation of the acids by the HIO_3 occurred. The oxidation in the presence of $AgClO_4$ was accompanied by the pptn. of almost all the I present as $AgIO_3$, when the $AgClO_4$ was added and also by the reduction of the $AgClO_4$ to finely divided Ag. The results show that only half the H_3PO_3 was oxidized. After the oxidation the soln. contained practically $H_2PO_4^-$ and very little $H_2PO_3^-$. H_3PO_3 only began to be formed after about half of the H_3PO_3 had been oxidized. At ordinary temp. the oxidation of H_3PO_3 to $H_2PO_4^-$ was the main reaction, the oxidation of H_3PO_3 alone was always slow. With $Hg(ClO_4)_2$ no HgI_2 was pptd. since it dissolved in the perchlorate formed. It is in formation. As with the Ag salts the Hg^{++} salts were the real oxidizing agents of the phosphorous acids. The Hg^{++} salts oxidized both H_3PO_3 and H_3PO_2 to $H_2PO_4^-$, while $AgClO_4$ only oxidized the H_3PO_3 to $H_2PO_4^-$. The oxidation by the Hg^{++} salt occurred at ordinary temp., the rate depending on the quantity of $Hg(ClO_4)_2$ present. According to the quantity of Hg^{++} salt added, either the Hg^{++} salt alone or the Hg^{++} salt and metallic Hg. were formed. Detailed descriptions of the courses of the reactions of solns. obtained and of ppts. formed are given, together with detailed equations representing the various oxidations.

EDWARD B. SANICAR

The reduction of soda by hydrogen. P. VILLARD. *Compt rend* 190, 1329-31 (1930) —A study was made of the action of H at high temp. upon free alkali, soda by preference because of the ease with which one can detect Na vapor. A current of H was passed over fused soda spread on the internal wall of a tube of extremely refractory glass or contained in a tube of Ag lightly closed at both ends to retain the soda and placed in the glass tube. The results were as follows: At 600° in the heated part of the tube carrying the soda the current of H was charged with vapor absorbing Na light very strongly. With a spectroscope and white light the absorbing vapor gave the characteristic spectrum of Na. It did not seem to be a question of disson. even at the higher temps. utilized (in the neighborhood of 600°) since the phenomena of the absorbing vapor disappeared entirely when H was replaced by N_2 .

E. H. CARTER

The system, thorium nitrate-ether-water between 0° and 20° . P. MISCIATELLI. *Gazz. chim. Ital* 60, 523-5 (1930) —It is necessary to have data on the quaternary system $UO_2 \cdot (NO_3)_2 \cdot Th(NO_3)_4 \cdot H_2O$, and therefore in turn of the ternary systems $UO_2 \cdot (NO_3)_2 \cdot Th(NO_3)_4 \cdot H_2O$ and $Th(NO_3)_4 \cdot Pt(NO_3)_2 \cdot H_2O$, to establish the conditions for the separation.

of Th from U in the presence of water by the aid of Et_2O (cf M, C A 23, 1551) Of the systems Et_2O , H_2O , $\text{Th}(\text{NO}_3)_4 \cdot \text{Et}_2\text{O}$ and $\text{Th}(\text{NO}_3)_4 \cdot \text{H}_2\text{O}$, only the latter is unknown, so this was first studied The following data give the temp and the no of g. of $\text{Th}(\text{NO}_3)_4$ in 100 cc of satd soln cryoscopic method —0.2, 1 —0.5, 2, —1, 5.2, —1.5, 9 —2.1, 13, —2.9, 16, —4, 20, —4.6, 23.5 —5.1, 26.4, —5.6, 27.3, —6, 33, —6.6, 37 —9, 41 —11.2, 43 —12.2, 44.5, —11.5, 46.2, —15, 47.7, —16.6, 49.8, —19.1, 51 —23.3, 53.1 —25, 55.7 —28.6, 58, —31.3, 59.2, —35, 60.6, —40.6, 62 —43.5, 64, synthetic method —43.5, 64 ($\text{Th}(\text{NO}_3)_4 \cdot 6\text{H}_2\text{O}$) analytic method —22.6, 12 0.65 10, 65.2 20, 65.6 Equil in mixts of $\text{Th}(\text{NO}_3)_4$, Et_2O and water was very slow in some cases 20 days being required The results of the measurements at 0° and 20° are given in tables and triangular diagrams, which must be consulted for the quant data The method used for obtaining these data is described in detail In the prepn of anhyd $\text{Th}(\text{NO}_3)_4$, the ordinary nitrate was freed of traces of Ce by pptn with $\text{Na}_2\text{S}_2\text{O}_3$ and H_2O_2 , crystn as $\text{Th}(\text{NO}_3)_4 \cdot \text{H}_2\text{O}$ over this to remove H_2O , and treatment of the residue with a current of dry air for several days

C. C. DAVIS

The system: uranyl nitrate-ether-water between 0° and 20°. P. MISCIATELLI *Gazz chim ital* 60, 839-42 (1930) —The binary systems Et_2O , H_2O , $\text{UO}_2(\text{NO}_3)_2 \cdot \text{Et}_2\text{O}$ and $\text{UO}_2(\text{NO}_3)_2 \cdot \text{H}_2\text{O}$ are already known Equil in the ternary system was much more rapid than with $\text{Th}(\text{NO}_3)_4$, Et_2O , H_2O A method of prepn of anhyd $\text{UO}_2(\text{NO}_3)_2$ was developed, viz, by distg its mixt with N_2O_5 and drying the residue in a current of dry air at 110° Incipient decompn of the anhyd salt is obvious from a change in color to orange yellow No basic salts are formed in this thermal decompn (cf Guempel, C. A. 24, 2367) The data for the isotherms at 0° and at 20° are given in tables and triangular diagrams, which must be consulted for the quant values Cf preceding abstr.

C. C. DAVIS

The structure of hydrogen phosphide and of hydrogen arsenide. G. NATTA AND E. CANAZZA *Gazz chim ital* 60, 851-9 (1930) —In previous expts on the structure of hydrides of elements of the 5th and 6th groups (cf N, *Giorn chim ind applicata* 12, 37 (1930)), analogies were found in the form and dimensions of the elementary cells of NiH_2 , PIH_2 , and AsH_2 , but this was limited to a comparison of the lattice const., and it was not proved whether a perfect isomorphism exists In the present paper, more refined methods have made it possible to settle this problem Of the binary compds. of H with elements of the 5th group, only NiH_2 has been studied by the x ray powder method (cf Mark and Pohland, C A 20, 130, de Smedt, C A 20, 1735), but the arrangement of the atoms in the mol was not established A new examn. of NiH_2 gave x ray photographs of extraordinary clearness, and these will be described in a later paper, with the detn of the structure The lattice const. at -170° was 5.08 A U., a value lower than that of the other experimenters, probably because of the thermal contraction at the lower temp SbH_2 and BiH_2 appeared to be isomorphous but of a different structure from NiH_2 , PIH_2 , and AsH_2 However, the weakness and lack of clarity of the lines leave this problem unsettled Good results were obtained with PIH_2 and AsH_2 PIH_2 was prepd from Ca phosphide and water and purified by drying with CaCl_2 , freezing out of heavier hydrides and sepn from H by solidification with liquid air It b -87.4° and m -132.5° AsH_2 was prepd from pure Zn and As_2O_3 in aq HCl , drying with CaCl_2 and freezing of the AsH_2 with liquid air Examn. of both PIH_2 and AsH_2 was carried out with an app. already used (cf N, C A 24, 4973) At approx 100° (abs) PIH_2 and AsH_2 are quite stable and do not oxidize in air, even when traces of other H phosphides or arsenides are present PIH_2 crystallizes in the cubic system Its elementary cell contains 4 mols, has a dimension of 6.31 ± 0.01 A U. at -170° and a vol of 251×10^{-24} cc. Its d is 0.896 AsH_2 is isomorphous with PIH_2 Its elementary cell contains 4 mols, with a dimension of 6.40 ± 0.02 A U. at -170° and a vol of 262×10^{-24} cc. Its d is 1.96 The position of the P and As atoms corresponds to a face-centered structure The x ray data did not permit the establishment of the position of the H atoms in the lattice From considerations based on the size of the constituent atoms, it is probable that both PIH_2 and AsH_2 belong to 1 of the 2 spatial groups Th^2 and Oh^2 The relations between these and the structure of NiH_2 will be the subject of a later paper

C. C. DAVIS

Salts of bivalent vanadium. JULIUS MEYER AND MARGOT AULICH *Z. anorg allgem Chem* 194, 278-92 (1930) —The soly of specially purified V_2O_5 in various concns of H_2SO_4 is tabulated If 3 g of V_2O_5 is dissolved in 30 g of 10% H_2SO_4 , the soln diluted and reduced with SO_2 , blue vanadyl sulfate is produced Electrolysis of the SO_2 -free soln. in an inert atm with a Pt, Pb or Hg cathode, 8-10 v and 2-3 amp causes further reduction Purple crystals of $\text{VSO}_4 \cdot 7\text{H}_2\text{O}$ (I) sep upon evapn. This product

is very unstable. The NH_3 , K and Pb complexes, e. g., $(\text{NH}_3)_3\text{V}(\text{SO}_3)_3 \cdot 6\text{H}_2\text{O}$, are more stable. Treatment of I with alkali salts of other acids gives only solns. Salts of org. acids decompose readily. Soln. of metallic V in mineral acids yields green vanadic solns. The metal as cathode or anode does not form V^{++} . B. A. SOTLE

Germanium. XXXVI. Extraction of germanium and gallium from germanite. L. W. I. PATNODE and R. W. WORM. *Ind Eng Chem* 23, 204-7 (1931).—About 1200 g. of germanite is chlorinated and the volatile chlorides are distd. The distillate includes the anhyd. chlorides of Ge, As, Si, S and Ga. These are subjected to fractional distn. This distillate contains HCl and GeCl_4 , which are sep'd by gravity. The GeCl_4 may then be hydrolyzed again to the dioxide. The yield is 85-90% of the Ge content of the ore. AMY LE VESCOMTE

The halides of the rare earths. IV. Samarium diiodide and the thermal decomposition of samarium triiodide. G. JANTSCH and N. SKALLA. *Z. anorg. allgem. Chem.* 193, 391-405 (1930).—Pure anhydrous SmI_2 is prepd. by treating a mixt. of 1 mol. of ordinary H_2O -contg. SmI_2 and 6 mols. of NH_3 with a current of H_2 , the temp. being gradually raised to 350° and finally to $600-650^\circ$ to remove NH_3 . After cooling the H_2 is replaced with N. To avoid the formation of basic salts the gases must be quite dry and free from O. SmI_2 so obtained can be reduced to SmI_2 by heating in a Au boat in a current of dry H_2 , the temp. being raised slowly in the course of 11 hrs. to $750-760^\circ$ and held at this temp. for 1.5 hrs. After cooling the tube is filled with pure N. SmI_2 is a deep green, almost black in color, and dissolves in H_2O to give a deep red color. The soln. evolves H_2 and is gradually decolorized with sepn. of a basic Sm^{III} salt. SmI_2 is more stable than the chloride but is vigorously decomposed by acids, even AcOH , with evolution of H_2 . The thermal decompn. of SmI_2 was carefully studied in a quartz tube under a pressure of 0.01 mm. of H_2 . The temp. rose to 500° in 4 to 5 min. at $550-570^\circ$, I began to come off, the reaction being very vigorous at $650-720^\circ$. Above 750° evolution of I was very slight. The material was heated at 850° for 45 min., then at 900° for 15 min., and then allowed to cool, an atm. of pure dry N being finally admitted. The product of this decompn. is SmI_2 . SmI_2 at 680° reacts as follows: $3\text{SmI}_2 \rightarrow \text{Sm} + 2\text{SmI}_3$. SmI_2 melts with decompn. at 840° and is extremely sensitive to H_2O , forming a red basic iodide in a few mins. in air. Cond. measurements verify the rapid decompn. of SmI_2 in H_2O soln. as follows: $2\text{SmI}_2 + 2\text{H}_2\text{O} \rightarrow 2\text{Sm}(\text{OH})\text{I}_2 + \text{H}_2$. If SmI_2 is dissolved in air free H_2O at 0° and treated with a cold satd. soln. of Na_2SO_4 , SmSO_4 is pptd. It is a bright, orange-red cryst. powder, slowly decomposed by H_2O and sol. in acids with decompn. H. STOEHRZ

Trichloromethyl perchlorate. L. BRICKENBACH and J. GOUTEAU. *Nachrichten aus der Chem.* 18, 520 (1930).—In presence of a little HCl , CCl_4 reacts with AgClO_4 at the ordinary temp., yielding trichloromethyl perchlorate CCl_3ClO_4 in 60-70% yield. By repeated fractionation in a vacuum in absence of moisture this can be obtained as a colorless liquid which solidifies at about -55° to colorless crystals. With water it yields HClO_4 and liberates I with violence. With alc. and other org. substances, e. g., tap grease, it reacts at the ordinary temp. with detonation. When heated it explodes. B. C. A.

The action of bivalent metals on persulfates of the alkali group. Synthesis of double salts of the type $\text{M}_2^{++} \text{M}^{++}(\text{SO}_4)_6$. OSSIAN ASCHIAN. *Z. anorg. allgem. Chem.* 194, 139-46 (1930).—Reduction of K, Na or Li persulfates with metallic Cu, Mg, Zn, Cd, Mn, Fe, Ni or Co gives a double salt, e. g., $\text{K}_2\text{Zn}(\text{SO}_4)_6 \cdot 6\text{H}_2\text{O}$. The reaction is practically quant. B. A. SOTLE

Alkylxyvanadylsalicylates. P. BRAUMAN. *Compt. rend.* 192, 156-7 (1931).— MeOH (or EtOH) + VOCl_3 + $\text{LiOC}_6\text{H}_4\text{CO}_2\text{Me}$ gives $\text{MeO}(\text{VO})\text{OC}_6\text{H}_4\text{CO}_2\text{Me}$ (or the EtO compd.). These are analogous to the Mo (Rosenheim) and Be (Berthel) compounds. *Z. anorg. Chem.* 34, 440 (1933) and the Cu and Ni compds. (Doak and Packer, *C. A.* 23, 829). V. F. HARRINGTON

Complex thiocyanates of quadrivalent molybdenum. G. A. BARRIÈRE. *Ann. chim. phys.* 12, 55-61 (1930).—The thiocyanates of quadrivalent Mo, previously unknown, have been prepd. by oxidizing the pyridine as well as the hexamethylenetetramine salts $(\text{C}_6\text{H}_5\text{N})_4\text{H}_2\text{Mo}(\text{CNS})_4$ and $(\text{C}_6\text{H}_5\text{N})_4\text{H}_2\text{Mo}(\text{CNS})_4$ with $\text{K}_4\text{Fe}(\text{CN})_6$. Both of the above salts are insol. in water; so the reaction is carried out by grinding each with $\text{K}_4\text{Fe}(\text{CN})_6$ in a mortar in the presence of water. With the $\text{C}_6\text{H}_5\text{N}$ salt, an intense blue-violet solid results, while with the tetramine, a green violet solid results. The analyses indicate that the products are $(\text{C}_6\text{H}_5\text{N})_4\text{H}_2\text{Mo}(\text{CNS})_4$ and $(\text{C}_6\text{H}_5\text{N})_4\text{H}_2\text{Mo}(\text{CNS})_4$. To prove that quadrivalent Mo is present, a method was developed by B.; the method depended on the observation that any Mo salt can be oxidized to Mo^{VI} by a cold soln. of ammoniacal $\text{Fe}(\text{CN})_6^{III}$, this being in turn reduced to $\text{Fe}(\text{CN})_6^{II}$. The equivs. of $\text{Fe}(\text{CN})_6^{III}$ reduced are a measure of the oxidation. The Mo was thus found to be

quadrivalent. The estn of $\text{Fe}^{\text{II}}(\text{CN})_6$ and $\text{Fe}^{\text{III}}(\text{CN})_6$ in solns of CN and SCN as well as halogens has already been worked out by B. C. A. 24, 4724. A. W. CONTI and

Studies on some complex chromiselenates. L. P. B. SARKAR AND S. N. BHATTACHARYA. *J. Indian Chem. Soc.* 7, 767-9 (1930).—The close analogy between the properties of complex chromisulfates and complex chromiselenates is discussed. The methods of prepn and analysis of chromoselenic acid, $\text{H}_2[\text{Cr}(\text{SeO}_4)_2]$, chromoselenomonosulfuric acid, $\text{H}_2[\text{Cr}_2(\text{SO}_4)(\text{SeO}_4)_2]$, chromoselenodisulfuric acid, $\text{H}_2[\text{Cr}_2(\text{SO}_4)_2(\text{SeO}_4)_2]$, chromoselenotrisulfuric acid, $\text{H}_4[\text{Cr}_2(\text{SeO}_4)_3(\text{SO}_4)_3]$, and the corresponding K salts are outlined. Na chromoselenic alum is described. L. L. QUILL

The system $\text{MgO-FeO-Fe}_2\text{O}_3$ in air at one atmosphere (RONFATS, MPAWIN) 8.

WOLF, LUDWIG. *Grundzüge der anorganischen Chemie*. 2nd ed., revised Leipzig: J. Deuticke. 230 pp. M 6, bound, M 8

7—ANALYTICAL CHEMISTRY

W. T. HALL

Differential potentiometric titration. IV. (a) Adaptation of the method to the use of hydrogen electrodes. (b) Test of standards for precise acidimetry. DUNCAN A. MACINNES AND IRVING A. COWPERTHWAIT. *J. Am. Chem. Soc.* 53, 655-62 (1931).—The way in which the differential electrometric method of titration can be adapted to the use of H electrodes is described in detail and 2 cuts show the necessary app. Expts with electrometric detn. of the end point by the method described, using const.-boiling HCl, K acid phthalate and benzoic acid as standards, gave results agreeing to within 0.01% or better. W. T. H.

Improved apparatus and method for the analysis of gas mixtures by combustion and absorption. MARTIN SHEPHERD. *Bur. Standards J. Research* 6, 121-67 (1931).—The app. described is of the Orsat type but contains many improvements. A buret is provided which prevents parallax errors, a device gives good illumination, a manometer-compensator is provided which is compact and less fragile than the usual type, the pressure balance being obtained by electric contact, a special stopcock insures accuracy; a water jacket is provided which does not require a very large rubber stopper, a distributor, or manifold, eliminates much capillary dead space, the absorption pipet works very rapidly, the combustion pipet eliminates the undesirable features of the common type, a sampling pipet is designed to permit manipulation without danger of contamination. Sources of leakage are avoided, the complete assembly is unusually convenient; the app. is supported entirely by metal, the assembly and replacement of parts is easily accomplished. Several procedures are given for making technical and accurate gas analyses. The app. was designed for the purpose of getting more accurate results and saving time. The original paper must be consulted for details. W. T. H.

Quantitative study of the boric acid-alcohol flame test. W. STAHL. *Acta Univ. Latviensis Kīm. Fakultat, Seriya 1*, No. 13-17 (in German 399-400) (in Lettish 369-99) (1930).—The H_2BO_3 esters give the green flame test when the reacting components are at about 100° below the b. p. of the EtOH or MeOH used. With rise in temp. the effect is increased. The concd. H_2SO_4 used in the test serves not only to liberate H_2BO_3 and remove H_2O but also has an effect upon the flame. The color can be estd. with the aid of a colorimetric scale. The best flame test is obtained when the temp. is not far from the b. p. of the alc. and the best reaction mixt. is 5 vol. MeOH to 1 vol. concd. H_2SO_4 . The dimensions of the app. have an effect upon the smallest quantity of H_2BO_3 that can be found but with the micro app. described in this paper, as little as 0.005 mg. H_2BO_3 can be detected. As much as 3% of water in the MeOH does not have any noticeable effect. W. T. H.

Spectrum analysis in assaying. A. A. FITCH. *Mining Mag.* 43, 81-5 (1930).—Procedure is outlined for the most valuable types of applications: (1) Qualitative test for a particular metal. This is recommended for sorting steel scrap. (2) Qualitative test for all metals present. This is valuable esp. in complete assays of Pt metals and rare-earth concentrates. (3) Approx. quantitative analysis—done by comparison with standard spectra. (4) More accurate quantitative detn.—done by measurement of intensity of lines by spectrophotometry. The accuracy of the method is about 5%. It is suitable for rapid analysis of steels where highest accuracy is not required.

Numerous other applications are suggested, the greatest advantage being rapidity. A BUTTS

Role of silica in the decomposition of fluorides. P FUCHS. *Chem Erde* 5, 99-105 (1930) —In the Berzelius method for the decomposition of insol fluorides, e.g., fluorspar, by fusion with alkali carbonate mixed with SiO_2 , the part played by the SiO_2 is in the formation of a double NaCa silicate . The carbonate and SiO_2 may be fused beforehand to avoid frothing in the actual analysis, the addition of NaOH is advantageous. B C A

Cobaltic sulfate as an oxidizing agent. SHERLOCK SWANN, JR., AND THEODORE S XANTHAKOS. *J Am Chem Soc* 53, 400-4 (1931) — $\text{Co}_2(\text{SO}_4)_3$ is best used as an analytical reagent for the quantitative oxidation of certain organic compounds to CO_2 . An electrolytic method is described for preparation of the reagent. B A SOUTER

The use of hydrogen sulfide and ammonium sulfide in analytical laboratories. R CATTELAU. *J pharm chim* [4], 12, 367-76 (1930) —Useful practical notes on the preparation of these reagents and precautions to be taken in preparation, washing and weighing the various metal sulfides are given. Their analytical properties are summarized in 5 tables. S. WALDBOTT

Organic reagents in qualitative inorganic analysis. B TOUGARINOFF. *Ann Soc Sci Bruxelles* 50, Ser B, 145-246 (1930) —The paper contains data concerning the use of about 100 different organic compounds which have been recommended for the qualitative testing of inorganic materials. A bibliography is given, also a statement concerning the nature of each reaction, the method of applying the test and the sensibility. Numerous reagents for the following cations are described: Al, Sb, Cu, Bi, Ba, Cd, Ca, Cr, Co, Sn, Fe, Mg, Hg, Mn, Ni, Pb, Sr and Zn. No less than 24 different tests are given for Cu, 13 for Fe, 10 for Mg, etc. After giving these data, the latter part of the paper states the results obtained in trying out some 15 of these tests—the alizarin test for Al, the rhodamine B, pyrogallol, benzidine and urotropine tests for Sb, the urotropine, dimethylglyoxime and rhodamine B tests for Bi, the antipyrine and K synthogenate tests for Co, the alizarin test for Mg, the benzidine test for Pb, the diphenylamine, diethylaniline and orange IV tests for Zn. The sensitiveness and specificity of each test are discussed critically and the 3 tests for Zn, the alizarin test for Al, the rhodamine B test for Sb and the dimethylglyoxime test for Bi are recommended especially. W T H

Determination of small quantities of hydrogen and oxygen in active charcoal. ALFRED STOCK, HERMANN LUX AND JOHN W R RAYNER. *Z anorg allgem Chem* 195, 188-63 (1931) —Degasify the charcoal and then burn it with a known volume of O_2 . Condense the resulting water, and from the quantity obtained compute the H content. Then, since the volume of CO_2 is equivalent to that of the O_2 , the residual gas, corrected for the O_2 in the water, gives the volume of O_2 from the O in the original coal. An apparatus is shown which is suitable for the analysis and the method is described in detail. W T H

Rapid colorimetric estimation of potassium. EARLE R CALBY. *J Am Chem Soc* 53, 539-45 (1931) cf C 1 24, 201 —Dissolve in 1 cc of water the dry residue from a solution and add 7.5 cc of a saturated solution of picric acid in 95% alc. After a precipitate begins to form in the stirred solution allow the mixture to stand at 20° for 45 min with stirring at 5-min intervals. Filter and wash the precipitate with ether. Dissolve the precipitate and compare the color with that obtained similarly from known quantities of K. W T H

Determination of magnesium in Portland cement and similar materials by the use of 8-hydroxyquinoline. J C REDMOND AND H A BRIGHT. *Bur Standards J Research* 6, 113-20 (1931) —Digest 0.5 g of cement with 20 ml of 6 N HCl , heating gently. Dilute to 150 ml and add Mg red indicator and coned NH_4OH until the solution is distinctly yellow. Boil 1-2 min allow the precipitate to settle and filter promptly. To the filtrate, add 1 ml of coned NH_4OH and 2 ml of 4% NH_4 oxalate solution. Boil 2-3 min, digest on the water bath for 30-60 min, filter and wash 5-6 times with hot water. Heat the filtrate (about 350 ml) to $60-70^\circ$ and add 20 ml of reagent (25 g 8-hydroxyquinoline dissolved in 60 ml of glacial AcOH and diluted to 2 l with cold water) and 4 ml of coned NH_4OH per 100 ml of solution. Stir mechanically for 10-15 min. Filter and wash the Mg precipitate with hot 0.37 N NH_4OH . Dissolve the precipitate in 50-75 ml of hot 1.2 N HCl , dilute to 200 ml, add 15 ml of coned HCl , cool to 25° and determine the hydroxyquinoline content by the KBr-KBrO_3 method of Berg. The results obtained with standard arylaceous limestone were excellent. The experienced analyst can determine MgO within 2 hrs by this method. W. T. H

Determination of magnesium as pyrophosphate. STANKO S MIHOLIC. *Bull Soc Chim Yoy Yougoslavie* 1, Pt 11 16-24 (1930) —To avoid the error due to rapid heat

ing of MgNH_4PO_4 , it is recommended to use Jena filtering crucibles and heat the ppt for 2 hrs at 450° in an elec. oven. W T II

Electrolytic separation of lead and antimony and its application to the determination of lead in tartar emetic. ELLA M. COLLIN AND HENRY J. S. SAND. *Analyst* 56, 90-3 (1931).—A study of the deposition potentials of Sb^{+++} , of Sb^{+} and of Pb^{++} in tartrate solns. showed the possibility of sepg Pb and Sb electrolytically from such solns. The method described depends upon the oxidation of the Sb to the higher state by I_2 in the presence of HCO_3^- . Tartaric acid is added and the soln made alk. with KOH. It is then electrolyzed by a current derived from an external source or from internal electrolysis. Dissolve 5 g. of tartar emetic in water add 3 g. of NaHCO_3 and oxidize the SbO with KI soln. Then add 4 g. KOH and 2-3 g. of tartaric acid. Dil to 300 cc. and electrolyze cold using a Pt cathode which has been coated with Cu. Rotate the anode. In 15-20 min quantities of 16-33 mg. of Pb can be detd. accurately. Directions are given for carrying out the method by internal electrolysis with a Zn anode and a 10% KCN soln in the anode compartment. W T II

Analysis of lead peroxide and of red lead. G. BRUNN. *Chem.-Ztg.* 55, 50-1 (1931).—Instead of using FeSO_4 or oxalic acid to reduce PbO_2 or Pb_2O_3 in the presence of acid, it is now recommended to use a soln. of $\text{Fe}(\text{NO}_3)_3$ and HNO_3 . A suitable soln. can be prepd. by treating FeSO_4 soln. with an excess of $\text{Pb}(\text{NO}_3)_2$ and allowing the ppt to settle. Add the soln. from a buret until the PbO_2 or Pb_2O_3 dissolves completely in the added HNO_3 and titrate the excess Fe with KMnO_4 . Excellent results were obtained in the 13 analyses cited. W T II

Analytical chemistry of rhenum. II. Determination of rhenum as nitron perchlorate after previous precipitation as sulfide. W. GERHARDT AND F. WEINKE. *Z. anorg. allgem. Chem.* 195, 283-300 (1931).—To ppt Re as sulfide, the soln. should be 3.6-4.2 N in HCl and the current of H_2S should be continued for at least an hr. It is well to start with the soln. hot and pass the gas for 2-2.5 hrs. while the soln. is cooling. Wash the ppt. with 0.3 N HCl which has been satd. with H_2S . Treat the sulfide with 3-5 cc. of 5% NaOH and with 2-3 cc. of perhydrol which converts the S to SO_4^{--} and the Re to ReO_4^- . After removing the excess peroxide by boiling neutralize with 2 N H_2SO_4 and ppt. the hot soln. with nitron. The oxidation of the sulfide can also be accomplished by fusion with Na_2O_2 which is particularly advantageous with sulfid. ppts. that have been dried and aged. Cf. C. A. 25, 50. W T II

Estimation of small quantities of tin. J. E. CLENNELL. *Mining Mag.* 43, 9-13 (1930).—The accuracy of the usual volumetric method of reduction of a stannic salt in HCl soln. with Fe or Ni, followed by titration with I_2 , was tested for detg. 1 to 30 mg. of Sn. Results with the smaller amts. were shown to be unsatisfactory. A colorimetric method that gives better results for 0.1 to 1 mg. of Sn is based on the blue color given by stannous salts with $(\text{NH}_4)_2\text{MoO}_4$. Several other reducing agents give the same color, so that the Sn must first be isolated. From the HCl soln. As and Sb are pptd. by Cu, which in turn is pptd. by Pb. The Sn is then pptd. on Zn from a slightly acid soln. This is dissolved in HCl and the soln. reduced by boiling with Pb foil. It is then poured into 10 cc. of alk. molybdate soln. (about 1% $(\text{NH}_4)_2\text{MoO}_4$ with 2% NaOH). A BUTTS

The lead reduction method for the determination of tin and the interference with it by copper and antimony. S. G. CLARKE. *Analyst* 56, 82-89 (1931).—The reduction of Sn from the quadrivalent to the bivalent conditions can be accomplished by boiling the soln. (1.2 N in HCl) with a strip of Pb foil for 1.5 hrs. in an atm. of CO_2 . The reduced soln. may be cooled and titrated with I_2 , the results being accurate if air is kept out by the stream of CO_2 during the entire expt. If, however, Cu is present in the soln., the results are low and the error is proportional to the Cu content. When Sb is present, the error is less but distinctly noticeable. The conclusion is drawn that the method should be applied only after the removal of Cu and Sb if accurate results are desired. W T II

The analysis of crude platinum and palladium. W. P. HORNE. *Mining Mag.* 43, 278-80 (1930).—Detailed procedure is given for analysis of material contg. 98 to 99.5% Pt or Pd, with detn. of the impurities and other precious metals present, as well as the Pt or Pd. A BUTTS

Rapid determination of zinc in brass. G. BRUNN. *Chem.-Ztg.* 55, 41 (1931). KARL ALTMANN-SBERGER. *JbJ* 41-2.—Thews and Harbison (C. A. 25, 661) recommended the approx. detn. of Zn by the color of the metal drillings, stating that more accurate results can be obtained by detg. the Cu and estg. the Zn by difference. B. points out that the KCNS- KI method is excellent for the rapid volumetric detn. of the Cu, and A. that the conventional iodide method gives excellent results. W T II

The analysis of wolframite and scheelite. C. STANSFIELD HITCHEN *Mining Mag* 43, 288-29 (1930) —The effect of W on the sepn of other elements makes a complete ore analysis difficult, and esp so in presence of Ta, Cl, Sn and Ti. The occurrence and mineralogy of W are described and analytical methods discussed at length, with recommendations and a diagrammatic scheme for procedure. A bibliography is also given. A. BUTTS

Sulfur-print tests. L. F. BENSON *Electrician* 106, 4-5 (1931) —The value of the S-print test, its simple operation and its sensitiveness are pointed out. The test can be applied to practically all steels of importance except when the Cr content is high, as in stainless steels and some non magnetic steels. W. H. BOYNTON

Rapid determination of carbon dioxide in carbonates. THEODOR HECIKO *Z. anorg. Chem* 44, 85-8 (1931) —Place the sample in a boat, cover with V_2O_5 , heat in a current of O_2 to 1200° and collect the resulting gas. After measuring the volume in a gas buret, absorb the CO_2 in KOH, the diminution in vol gives the CO_2 . W. T. H.

Improved method for determining carbon dioxide. C. A. JACOBSON AND JOHN W. HAUGHT *Bull West Va Univ Sci Assoc* 2, No 4, 8-13 (1930); cf *C A* 24, 4237 —The carbonate is decompd. with 20 cc. of 15% $HClO_4$, the evolved gas escapes through a condenser into a bubble tube, is dried with "dehydrite" and the CO_2 absorbed in ascarite. The results obtained in over 40 expts are reported. W. T. H.

Assay of calcium carbide and Czechoslovakian standards. R. VONDRACEK *Chem. Listy* 24, 191-7 (1930) —The provisional Czechoslovakian standards make use of a modified Caro's app for the assay of CaC_2 , this method is not exact, and an app giving more accurate results is described. The yield of C_2H_2 should not be given in cc. of moist gas, as the Gay Lussac formula is only applicable to dry gases. The exptl error in the assay of carbide should not exceed 2%. B. C. A.

Determination of free acidity of chrome alum. J. R. GUARCI *Afinidad* 10, 157-8 (1930) *Chimie & Industrie* 24, 1082 (1930) —The various alums (K, Fe, Cr) obtained as by products in several mfg processes cannot generally be used directly on account of their free H_2SO_4 content, which must be neutralized, and therefore detd., before use. Bellucci and Lucchesia method, which is used in $Al_2(SO_4)_3$ plants, is based on the colorimetric reaction of hydrazylized salts with methyl orange, and cannot be used for Cr alum on account of the color of the Cr ion. A suitable method is based on the insolv. of Cr alum in alc. A 15% soln of the alum is treated with 3.5 times its vol of 95% alc., allowed to stand 24 hrs., filtered and titrated with 0.2 N K_2CO_3 in the presence of phenolphthalein. A. PAPINBAU COUTURE

Colorimetric determination of the sulfate ion in water, coal, etc. P. GUARNIERI *Ind. Ital. Cons. Alim.* 3, 161 (1930), *Chimie & Industrie* 24, 814 —The method consists in treating 100 cc. of the water at the boiling temp. with an HCl soln of $NaCrO_4$, neutralizing with $CaCO_3$, filtering the ppt. and comparing the color of the filtrate with that of standards. A. P. C.

Determination of iodide in the presence of other halides. HUGO DITZ *Z. anorg. allgem. Chem.* 194, 147-50 (1930) —The statement of Gorbachev and Kasat. Lima (*C A* 24, 4210) that IO_3^- cannot be used as an oxidizing agent for the detn. of I^- in the presence of other halogens is refuted. Early expts. of Ditz and Margosches (1901) led to the conclusion that it is possible to have a simple and exact process for the detn. of I^- with Br^- and Cl^- on the different behavior of the halides toward IO_3^- with definite H^+ concns. Further investigations (1904) led to a process for the detn. of I^- by which a definite excess of KIO_3 soln and a definite amt. of H_2SO_4 are added to the I^- soln. The liberated I_2 is sepd. and titrated with $Na_2S_2O_3$ soln. A. S. S.

A new microdetermination of the calcium ion and of phosphorus. M. MOUSSEAU AND (MILIE) N. BLISSON *Bull. pharm. Sud-Est, Schweiz. Apoth. Ztg* 68, 654-8 (1930), cf (*C A* 24, 4731) —To det. both Ca^{++} and PO_4^{--} in a single sample, ppt. the P as $FePO_4$ by adding 1 cc. HCl and carrying out a basic acetate sepn. Dissolve the ppt. in HNO_3 and det. P by the molybdate method. Take the filtrate and ppt. Ca with Na tungstate. S. WILDEROTT

Determination of small quantities of volatile organic acids in sulfuric acid solutions. D. N. CRAIG *Bur. Standards J. Research* 6, 109-92 (1931) —In the study of the corrosive effect of various org. substances on the plates of Pb accumulators, it was found that sometimes small quantities of AcOH in the electrolyte caused corrosion of the + plates. In this connection, as well as in the study of the adequacy of treatment of wood separators, it became necessary to det. small quantities of AcOH in the presence of 30-40% H_2SO_4 . It was also desirable to test for formic acid. Fifty ml. samples of the acid were treated with sufficient 20% NaOH soln. to neutralize about 35 ml.

of the acid. The mixt was carefully evaporated to dryness by a bath kept at 65–75° until most of the liquid was evaporated and then heated about 10° higher. Water was added and the evaporation repeated to make certain that all volatile acid was removed. The apparatus used for the distn was ingeniously devised so that there was no danger of any spattered acid getting into the distillate. The total acidity of the distillate was detd. The neutralized soln was then treated with Na_2CO_3 and KMnO_4 to destroy the formic acid and then another distn with H_2SO_4 served to remove AcOH alone. The acidity of the distillates was detd by electrometric titration using the differential method of MacInnis and Dole somewhat modified. The values obtained were also compared with those obtained by indicator titrations using phenolphthalein. The potentiometric method was found excellent for titration solns of low acidity as obtained from electrolytes contg about 0.1% of AcOH or formic acid in 40% H_2SO_4 . W. T. H.

Microdetermination of urea by Nicloux and Welter's method. R. G. HENRY AND P. GILAZ. *Compt rend soc bio* 101, 726–8 (1929). —The rapidity and precision of the method are increased by pptg the serum with Tanret's reagent, treating the filtrate with 0.2 cc of a 5% soln of anthrylrol in MeOH (or 1% in AcOH), and collecting the ppt on a filter. After washing finally with satd alc dianthylethylamine, the ppt is dried in a current of air at 125° for 5 min. H. C. A.

A new method for the conversion of alkali sulfate to alkali chloride for use in silicate analysis (MVLCA) 6. Calorimetric method for determining silicon (KINO) 11B. Density of H_2SO_4 solutions of CuSO_4 (CHAPAS, *et al*) 18.

CURTMAN, LOUIS J. A Course in Qualitative Chemical Analysis. New York: The Macmillan Co.

YNGLEDER, CARL J. Laboratory Record Book in Qualitative Analysis. 110 pp. Laboratory Record Book in Quantitative Analysis. 120 pp. New York: John Wiley & Sons, Inc. \$1 each.

FRANÇOIS, MAURICE. Manipulations de chimie analytique appliquée, analyse des produits pharmaceutiques, chimiques et galéniques, analyses des matières alimentaires, hydrologie et analyse des eaux, analyses biologiques et toxicologiques. Paris: Le François. 372 pp. P. 60.

TALBOT, HERBY J.: Quantitative Chemical Analysis. Revised by L. P. Hamilton and S. G. Simpson. New York: The Macmillan Co.

Analyzing gases by absorption. I. G. PARRINELLO. A. C. Ger 513,492, Feb. 10, 1927. Manipulative features and app. are described.

8—MINERALOGICAL AND GEOLOGICAL CHEMISTRY

EDGAR T. WHERRY AND J. V. SCHAIKTS

Optical data for some rare minerals. T. BARTH AND H. BERMAN. *Chem Zvest* 5, 22–12 (1930). —Mention is made of various media of high n and of the dispersion method for detg under the microscope the ns of immersed mineral particles. Optical data are given for some 30 minerals. H. C. A.

Algodonite and whitneyite. F. MACHATSCHKE. *Neues Jahrb Mineral Geol.* Abt. A, Beilage Bd. 59, 137–148 (1929). —The minerals are crystallographically isomorphous. The regular crystals, apparently metallic Cu contg As, have a 3.617–3.651 Å U. The hexagonal crystals contain less Cu than corresponding with the formula Cu_2As_2 and have a 2.599, 2.598, c 4.215, 4.213 Å U., with d 8.71, 8.72. H. C. A.

The system $\text{MgO-FeO-Fe}_2\text{O}_3$ in air at one atmosphere. H. S. ROBERTS AND H. E. MERWIN. *Am J Sci* [5], 21, 115–57 (1931). —The system $\text{MgO-FeO-Fe}_2\text{O}_3$ has been investigated at temps above 1000° on the isobar $\frac{1}{2}$ atm O_2 . In the binary system $\text{MgO-Fe}_2\text{O}_3$ the only compd. is $\text{MgO-Fe}_2\text{O}_3$ which dissociates slightly on heating and begins to melt at 1750 ± 25°. Solid soln extends for a short distance toward MgO . In the binary system $\text{FeO-Fe}_2\text{O}_3$ the stable phase at temps below 1386 ± 5° is a hematite solid soln contg less O than Fe_2O_3 ; just above 1386° the stable phase is a magnetite solid soln contg considerably more O than Fe_2O_3 . The ternary system contains 2 solid soln fields. The first begins at MgO below 1000° and extends with increasing temp toward $\text{MgO-Fe}_2\text{O}_3$ and FeO , a liquid phase appears at 1770 ± 25°, where the solid contains Fe oxide equiv. to 73% Fe_2O_3 . The second solid soln extends with increasing temp. from $\text{MgO-Fe}_2\text{O}_3$ toward the Fe oxide boundary, which it reaches

at 1386 \pm 5°, solid solns of about 15% MgO in MgO Fe₂O₃ are possible from 1750 to 1000° or below ALDEN H. EMERY

The genesis of lodestone. MAAR C. BANDY. *Econ Geol* 25, 871-5(1930).—B discusses recent articles by Newhouse (*C. A.* 23, 4649) and Gruner (*C. A.* 23, 4167), on the origin of lodestone. The theory proposed by N. that lodestone is due to a brownish oxidized magnetite, does not meet all the conditions under which it occurs naturally and is formed artificially. The theory that it is formed by the magnetic field of the earth is equally inadequate. Lightning striking outcrops of magnetite ore is suggested as a theory for the origin of lodestone. ALICE W. EPPERSON

A contribution to the study of moonstone from Ceylon and other areas and of the stability relations of the alkali feldspars. EDMUNDSON SPENCER. *Mineralog Mag* 22, 291-367(1930).—Moonstone feldspar from pegmatite near Ambalangoda was found to possess a micropertthite structure. The lamellas (or ab) have sep_d mainly along the plane Δ (13 0 2) which is also the plane of schiller. A coarser but fainter perthite occurs, which sep_d at an earlier stage along the prism planes 110 and 320. Schillerized feldspars from Kandy, Burma and Colorado were found to be similar. The schiller color is dependent on the size and abundance of micropertthite lamellas and is mainly due to the reflection and scattering of light at the boundaries of these. The sp gr, optical axial angle and extinction angles are dependent on the proportions of the 2 feldspars present. The schiller color and micropertthite structure can be destroyed by heat treatment, and there is a reduction in sp gr, optical axial angle and extinction angle. The changes appear to be due to a re-soln of the (or, ab) members into each other to form the original (Or, Ab) phases. Structures can be developed on cleavage faces where there is little evidence of micropertthite by heating with water and CO₂ under pressure, causing selective decompn. The phenomena observed on heating indicate that there are 2 solid soln. changes, at (a) 900-1000° between a single solid soln phase and the 2 (Or, Ab) phases, and (b) 500-700° for the (or, ab) sep_d either from the 2 (Or, Ab) phases or from one of them. Structure in the Ambalangoda microcline-micropertthite indicates that the conversion to microcline evidently took place prior to the exsoln. of the micropertthite and hence probably above 700°. It undergoes none of the changes in properties found for moonstone feldspar and no re-soln takes place on heating. Microcline appears to be the stable form of K feldspar and to be incapable of holding albite in solid soln up to temps near its m p. A. M. B.

Mineralogy of Western Australia. EDWARD S. SIMPSON. *J. Roy. Soc. W. Australia* 15, 99-113(1928-29), cf *C. A.* 24, 3969.—Descriptions and in most cases chemical analyses are given for: apatite, gray and white beryl (contg. CaO 0.72 and 0.92%, resp.), cinnabar, glaucophane, siderite and corundum, manganilmenite and menacanthite, manganocolumbite, Ta₂O₅ 16-48, Cr₂O₃ 62-34%, resp. microcline, Ta₂O₅ 77.00, Cr₂O₃ 3.64, spinel, tourmaline, both schorl and dravite and vesuvianite.

W. L. HILL

Minerals from the Adamello Mountains (Trentino). C. GOTTFRIED. *Chem. Erde* 5, 106-12(1930), cf *C. A.* 24, 4243.—The minerals described are from marble (metamorphic limestone) at contact with tonalite. Garnet as reddish brown dodecahedra, with d 3.512 and n_{max} 1.7585 gave analysis I, corresponding with grossular 80.55, andradite 8.13, sillimanite 4.48, pyrope 6.85%. Vesuvianite as olive green crystals with $a = 1.05391$ d 3.336 n 1.719, n 1.715, gave analysis II, agreeing with Tschermak's formula $\text{Si}_2\text{Al}_2\text{Ca}_2\text{Mg}_2\text{H}_2\text{O}_7$. X-ray rotation photographs gave a unit cell of dimensions a 22.03 c 11.89 Å U, contg. 23 such mols., z e. 168 O atoms. The space group is D_{4h}^4 .

	SiO ₂	TiO ₂	Al ₂ O ₃	Fe ₂ O ₃	FeO	MnO	MgO	CaO	H ₂ O
I	39.66	0.46	18.77	2.52	1.92	0.07	2.06	34.79	0.07
II	37.03	0.75	16.25	3.13	1.00	0.04	2.45	36.39	2.29

II also ZnO 0.29, Na₂O 0.28, K₂O 0.03, F 0.31, total less O for F 100.15%

B. C. A.

Analytical determination of the isomorphous variation in rock-forming minerals. H. VON PHILIPSBORN. *Chem. Erde* 5, 233-53(1930).—The powd. rock, a pyroxene-granulite from Hartmannsdorf, Saxony, was sep_d into different fractions with Clerici's soln and an electromagnet. Analysis Ia is of the heavier and darker fraction (d 3.55-3.58, n [$= 1/2(\alpha + \gamma)$] 1.715) of the hypersthene and fb of the lighter and paler fraction (d 3.49-3.55, n 1.706). IIa of the heavier fraction (d 3.40-3.42, n 1.708) and IIb of the lighter fraction (d 3.37-3.40, n 1.698) of the monoclinic pyroxene. Analysis III is of the heavier fraction (d 2.69-2.71, n 1.555-1.560) of the plagioclase, correspond

ing with $Ab_{17}An_{83}$, the lightest fraction with $n = 1.549$ corresponds with $Ab_{18}An_{82}$. Analysis IV is of the garnet ($d = 3.92$, $n_x = 1.766$) and V is ilmenite. From optical data on the still lighter and heavier fractions of the pyroxenes their composition is arrived at by extrapolation. VI is the bulk composition of the rock ($d = 3.151$) corresponding with hypersthene 28.53, monoclinic pyroxene 29.60, plagioclase, 32.59, orthoclase 0.41, garnet 3.31, ilmenite 4.35, magnetite 0.35, apatite 0.53%.

	SiO ₂	TiO ₂	Al ₂ O ₃	FeO ₂	FeO	MnO	MgO	CaO	Na ₂ O	K ₂ O	Total
Ia	49.42	0.54	2.64	3.42	23.15	0.26	18.17	0.91	1.23	0.19	99.96
Ib	50.30	0.53	3.26	3.26	20.09	0.26	20.14	1.11	1.28	0.22	100.75
IIa	49.50	0.69	4.39	2.81	9.90	0.21	11.87	19.31	1.47	0.16	100.37
IIb	50.12	0.65	4.74	2.01	8.57	0.24	12.87	19.84	1.60	0.14	100.78
III	55.41	—	29.42	0.40	—	—	—	11.11	n d	n d	96.34
IV	38.50	0.88	21.77	1.47	19.61	0.72	8.75	8.50	—	—	100.23
V	—	48.3	—	10.3	38.3	0.3	1.7	—	—	—	100.4
VI	48.61	2.45	12.33	3.22	9.64	0.17	10.12	10.37	2.02	0.18	100.47

V, also insol silicates 1.5, VI, also P_2O_5 0.36, Cr_2O_3 0.10%.

B. C. A.

Chemical formula of the zirconium pyroxenes and zirconium pectolite. W. H. ZACHARIASEN *Norsk Geol. Tids.* 11, 216-8(1930).—From a study of analytical data obtained from the Zr pyroxenes, livenite, wöhlerite, and hiortdahlite and the Zr pectolite, rosenbuschite, it was found that the general formula $R_2Si_2X_2$ held for all 4 minerals where R represented all cations save Si, and X the anions. This proves that the names "Zr pyroxenes" and "Zr pectolite" have no chemical justification. H. H. MOSHER.

Sapphirin from Alpe Brasciadega, Val Codera, Italy. H. P. CORNELIUS AND B. DITTLER, *Neues Jahrb. Mineral. Geol., Abt. A, Beilage* Bd. 59, 27-64(1929).—The sapphirin, isolated by means of H_2SO_4 and HF contained SiO_2 15.19, TiO_2 0.25, Al_2O_3 61.69, FeO 4.31, MnO 0.12, CaO 0.49, MgO 16.23, $H_2O + 16O, H_2O = 0.10\%$, corresponding with the formula $Mg(Fe, Mn, Ca, H)_3Al_6Si_6O_{21}$. B. C. A.

Isomorphous series in the tourmaline group and the genetic relationships between tourmaline and the micas. W. KUNITZ, *Chem. Erde* 4, 208-51(1929).—Theoretical. B. C. A.

Tourmaline group. F. MACHATSCHKE, *Chem. Erde* 4, 455-7(1930).—A criticism of the formula of tourmaline proposed by Kunitz (preceding abstract). B. C. A.

Thuringite from Schmiedefeld, Thuringia. H. JUNO AND E. KÖHLER, *Chem. Erde* 5, 182-200(1930).—Olive-green material consisting of a compact aggregate of fine scales with $d = 3.187$ gave on analysis SiO_2 20.82, TiO_2 trace, Al_2O_3 17.64, Fe_2O_3 8.70, FeO 37.96, MgO 4.15, $H_2O + 1031, H_2O = 0.07$, total 99.65%. The analysis agrees with the formula $14(Fe, Mg)_2O \cdot 5(Al, Fe)_2O_3 \cdot 12H_2O \cdot 7SiO_2$. Most of the water is expelled at about 420° , and the dehydration curve is very similar to that of kaolin. X-ray powder photographs of thuringite have been compared with those of quartz, magnetite and kaolin, some of the magnetite and kaolin lines correspond, but those of quartz are absent. Magnetite has a unit cube of edge 8.393 Å. U. contg. 8 mols. B. C. A.

Studies on the zeolites. I. General review. MAX H. HEY, *Mineralog. Mag.* 22, 422-37(1930).—A review of the general characteristics of the zeolites and their relations to other groups of minerals and to certain artificial products, preparatory to a detailed examination of the various species. A bibliography is given. A. M. BRANT.

Steatite from the Münchberg gneiss area. F. DEUBEL, *Chem. Erde* 5, 87-95(1930).—A new occurrence of steatite has been found near Schwarzenbach a. d. Saale on the southeast border of the Münchberg gneiss area in Bavaria. It is shown to be a replacement of quartzite, no doubt by hydrothermal action. Analysis of the material gave: SiO_2 61.95, Al_2O_3 0.87, FeO 0.57, MgO 32.16, loss on ignition 4.99, total 100.54%. B. C. A.

Dehydration and rehydration of kaolin. P. SCHLACHTSCHABEL, *Chem. Erde* 4, 395-419(1930).—Kaolin after being heated at $400-800^\circ$ is capable of taking up H_2O again. This is effected very slowly at 110° , but under pressure at $175-205^\circ$ all is restored after 100 hrs. The rehydrated kaolin differs from the original material by being partly sol. in HCl , in its dehydration curve, x and x-ray pattern, but these differences become less marked after prolonged heating of the material in H_2O at 200° . B. C. A.

Halloysite from Elbingerode [Harz]. O. H. ERDMANN-SDORFFER, *Chem. Erde* 5, 90-8(1930).—Nodules of grayish green halloysite (analysis 1) coated with white powdery material (analysis 11) occur in limonite in the "Grosse Graben," which has been worked for Fe and Mn ores and pyrite. Both are optically isotropic with $n = 1.515-1.528$.

It gives the ratios $\text{SiO}_2/\text{Al}_2\text{O}_3/\text{H}_2\text{O} = 2.20/1.490$. Between 103° and 470° very little water is lost and the material then behaves like kaolin. An x-ray powder photograph shows a few faint lines coinciding with those of kaolin. The ignited material gives no x-ray spectrum.

	SiO_2	Al_2O_3	FeO	Fe_2O_3	MgO	CaO	Na_2O	K_2O	H_2O (100°)	H_2O (470°)	Total
I	38.27	23.53	5.31	0.10	0.02	0.23	1.60	0.64	14.19	12.70	100.99
II	29.00	30.07	2.08	0.01	0.04	0.40	1.31	1.24	12.70	15.15	101.09

Also traces of P_2O_5 , TiO_2 , SO_3 , CO_2 , MnO B C A

Nontzonite. W. NOLL. *Chem. Erde* 5, 373-84(1933). —Weinschenk (Z. Krist. *Mitt.* 23, 153-64(1907)) regarded nontzonite as analogous to kaolin. He gave the formula as $\text{H}_2\text{Fe}_2\text{Si}_2\text{O}_8$, but very few analyses agree with such a formula. Dehydration curves and x-ray powder photographs also show no relation to kaolin. The water is "reducible." A new analysis of green greasy nontzonite from Tachau, Bavaria, gave "reducible" SiO_2 43.02, Al_2O_3 10.55, Fe_2O_3 2.15, FeO 1.37, H_2O 21.53, total 99.93%, corresponding with the ratios $\text{R}_2\text{O}_3/\text{SiO}_2/\text{H}_2\text{O} = 1.249/4.53$ B C A

The chemical composition and crystallographic and optical properties of a Vesuvian Libitonsite. GUSTO CARONNI. *Atti III congresso sci. chim. pavia aprile 1930*, 333-7. —A blue Vesuvian libitonsite of the formula $(\text{Cu}, \text{Na}, \text{K})_2\text{Si}_2\text{O}_6$, m. p. 780° , was sep'd. into Cu-poor and Cu-rich fractions by flotation in a heavy liquid. The former contained blue monoclinic crystals with $d = 2.45-2.57$, $n_x = 1.536$, $n_y = 1.5307$, $n_z = 1.145$. The latter had $d = 2.74$, was optically —, and had the plane of the optic axis perpendicular to the plane of symmetry. The sp were 1.545 and 1.574. Chem. analyses of both fractions are given. J B AUSTIN

Blomstrandite from Kabuland, HARALD BJØRVEKE. *Norsk Fels II*, 330-9(1930). —This deposit of possible industrial importance was found in gang chiefly composed of xenotime, orthite, feldspar, quartz and magnetite. Two crystal forms were noted, the typical blomstrandite and a eucristite-like type. The mineral yielded upon analysis Ta_2O_5 19.11, TiO_2 31.84, UO_2 3.78, ThO_2 1.46, rare earth oxides 35.19, FeO 4.30, CaO 0.88, MgO 0.38, ignition loss 3.70%. By spectro-analysis Y, Th, Hf, Cp, Yb, Tu, Er, Ho, Dy, Tb, Gd, Sm and Nd were detected. H H MONTGOMERY

Chemical and spectrographic investigations on an apatite in Chinese strata of recent origin. GUSTO CARONNI. *Atti III congresso sci. chim. pavia aprile 1930*, 333-40. —A specimen of apatite from Tsungtao was examined. It was a solid solution of fluorapatite, chlorapatite, hydroxapatite and oxapatite. The crystals, which were badly formed, appeared to be unusual, —, with $d = 2.23$, $n_x = 1.52$, $n_y = 1.50$. Spectroscopic examination showed no Ce but the presence of La, Sm, Eu, Er and Yt. A chem. analysis is given. J B AUSTIN

Copiferous melanterite from the Skouroussa mine, Cyprus. MAX H. HEY. *Mining Mag.* 22, 413-8, 10. —A rather large greenish-blu crystals of melanterite was found in an ancient carved type in the ore body near Lefka, Nicosia district, Cyprus. The crystal habit is unusual, being tabular on $b(010)$. Partial analysis gave CuO 2.15%. The crystallographic measurements are given. A. M. BRANT

A suspected meteoric specimen from northern Arizona. L. F. BRADY. *Amer. J. Sci.* 13, 21, 170-174. —The specimen analyzed: metallic Fe 57.41%, Fe as FeCl_2 0.80%, Fe as FeO 0.28%, S 0.073, S 0.432, Mn 0.120, Cl as FeCl_2 1.300, P 1.110, Ni 4.040, Cu 0.047, C (graphite) 45.530, combined C 2.202, O as FeO 0.943 and insolubles (mainly FeO and SiO_2) 1.700%. The outer surface resembles weathered hematite. In section it appears to be an intergrowth of crystalline Fe in a ground mass of graphite. Etched surfaces show very fine Neumann lines. Its sp gr is 3.7. A. H. E.

A new iron meteorite from Pojoaque, New Mexico. L. F. BRADY. *Amer. J. Sci.* 13, 21, 174(1931). —The specimen is of Ni-Fe. There are several spherical inclusions of chrysolite which seem to be surrounded by a layer of kamacite. Its sp gr is 6.2. ALLEN H. EMMETT

New observations on the tectites of Indochina. A. LACROIX. *Compt. rend.* 191, 893-9(1930). —C 4 23, 2335-6. —The tectites in Indochina are widely disseminated. In order to consider the uniformity of their chemical composition, tables of results of analyses of tectites taken from widely separated places are given. These show very close agreement. Various theories of origin are discussed, and the only one considered not impossible is that of cosmic origin, as meteorites. Indications defending this hypothesis are cited and discussed. ALICE W. EPPERSON

Synthetic sulfide replacement of ore minerals. JAMES C. RAY. *Econ. Geol.* 25, 433-51(1930). —Geochem. expts. were carried on with the object of determining the point at

which dispersion of bornite in chalcocite occurs as well as such replacement phenomena as might be induced under lab conditions. Tabulated data are given, as are results of microscopic examn of the repolished surfaces of the treated ore fragments. The expts show that mixed sulfides in the presence of H_2O only can generate their own conditions to bring about changes in these sulfides without addn of m tals to the local system in which the changes are taking place. R advances the hypothesis of "deep-seated supergene alteration" for this type of massive deep-seated bornite, and believes the principles here advanced have a general application to Cu deposits of hydrothermal origin although the discussion is based specifically on the Butte ores. A. W. I.

The texture and origin of some banded or schistose sulfide ores. W. H. N. R. W. HOUSE AND G. F. FLAHERTY. *Econ Geol* 25, 600-20(1930).—Conclusions. Criteria of flowage and plastic deformation of an ore mass, based on apparent flow structures as seen in hand specimens, are usually unreliable. An ore mass composed in whole or in part of galena, sphalerite and chalcopyrite, which has been strongly d-formed, will show in its texture that flowage has taken place unless complete recrystn has occurred. Extent of recrystn and size of crystals are a function of strain and compn in the crystal and of time and temp. Unless complete recrystn of the entire ore mass takes place 2 generations of each mineral should be found. Quartz is an introduced mineral in schist replacement deposits, showing no preferred direction of elongation or orientation. Ores are particularly discussed occurring at Blue Hill Me., Milan, N. H., Mandy, Manitoba and Rammelsberg, Goslar, Germany. ALICE W. I. PIERSON.

Pseudo-eutectic textures. G. M. SCHWARTZ AND CHARLES F. PARK, JR. *Econ Geol* 25, 638-63(1930).—S and P discuss Lindgren's paper (C. A. 25, 205) and describe a specimen of ore contg covellite intergrowths, which furnishes addnl proof that graphic intergrowths may form by replacement. It is considered that pseudo-eutectic textures develop by replacement, even if supergene, and they may also develop by more or less contemporaneous crystn as in ilmenite and silicates, though ilmenite could not form a eutectic with several minerals in the same rock. It is possible for the texture to develop in hard minerals, but it is relatively uncommon and has not been observed except in gabbro and titaniferous magnetites. A. W. I.

A colloidal origin of some of the Kennecott ore minerals. S. G. LASKY. *Econ Geol* 25, 737-57(1930).—The sulfides that make up Kennecott ore bodies were transported in highly dispersed condition and were flocculated on reaching such an ideal flocculating agent as limestone. Alumina along thin gouge seams also served to flocculate sulfides as well as to dam back the solns. Chalcocite and S were always present. The chalcocite, in addn to the covellite that it was able to dissolve from the surrounding material at high temp., contained in soln also a certain amt. of covellite formed by reaction with excess S, which at this stage was present in relatively small quantity. With lowering of the temp., the covellite tended to unmix and did so in part until equil was reached. The equil pt. lies apparently above 85% CuS. A. W. I.

The incipient oxidation of galena. ALFRED L. ANDERSON. *Econ Geol* 25, 528-42 (1930).—Study of incipient states of alteration of galena indicates it to be more generally oxidized by air-water processes than by solns of ferric sulfate. Much of the limonite represents a later addn from external sources. The first product obtained by oxidation of galena is always anglesite. With admixt of pyrite and chalcopyrite, the reactions become more complex. Because of formation of electrolytic cells between various sulfides in contact with the oxidizing solns, the attack on galena is greatly aided, while the oxidation of pyrite or chalcopyrite is retarded. On this basis galena oxidizes earlier and with more ease than most other sulfides, except sphalerite. Galena remains in the oxidized croppings only because the coating of insol alteration products prevents ready access by fresh oxidizing solns. Covellite is a common but not abundant mineral formed by replacement of galena. Four main types of structures may be recognized: (1) regular borders, (2) botryoidal structures, (3) stranded borders and (4) platy structures, each referring to the relation the covellite bears to the galena. ALICE W. I. PIERSON.

Limonite types derived from bornite and tetrahedrite. ROLAND BLANCHARD AND P. F. BOSWELL. *Econ Geol* 25, 557-80(1930).—Neither bornite nor tetrahedrite contains sufficient S to dissolve itself completely. Field evidence has not established definitely the proportions of pyrite needed to effect complete oxidation of bornite and tetrahedrite in nature. Hydrolysis of $Fe_2(SO_4)_3$ during oxidation liberates acid and may reduce below the theoretical ratio the pyrite needed for complete oxidation of bornite and tetrahedrite. Field evidence shows that an overwhelming proportion of pyrite is needed to prevent formation of indigenous limonite after bornite, and such proportion is not usual. Five limonite types derived from oxidation of bornite have

been identified and described, a triangular boxwork being most distinctive. Leached outcrops derived from bornite generally carry a "celadon" limonite—a characteristic product, orange colored, with soft texture on fresh surfaces. Two limonite types derived from oxidation of tetrahedrite have been identified and described. That called contour boxwork is more distinctive. Leached outcrops from tetrahedrite usually carry local incrustations of Fe oxides that assist in identification. A. W. I.

A mother lode gold ore. CARLTON D. HULIN. *Econ Geol.* 25, 348-53 (1930).—H. describes specimens from the Kennedy Mine, one of the largest of the Au mines operating on the Mother Lode, in Amador County, Cal. The relative degree of favorability of the several host minerals for Au as observed in these ores is: chalcocite, pyrite, galena, apatite, carbonates, pyrite, sphalerite, and quartz, the first 5 minerals are about equally favorable, and quartz is the least favorable. A. W. E.

Geology of the Panamint silver district, California. F. MAC MURPHY. *Econ Geol.* 25, 305-24 (1930).—The paper includes a history of mining in the district, its location and topography, a description of the rocks and their structure, and a discussion of the ore deposits. The occurrence of numerous clean-cut Ag-bearing quartz veins is described. They outcrop in limestone, schist or slate, but those in limestone are more numerous and persistent. The principal and almost exclusive vein filling is milk-white, sometimes glassy, quartz of coarsely crystalline massive texture. Ag is almost wholly contained chemically bound in tetrahedrite, the principal sulfide mineral. Oxidation and supergene enrichment products are not abundant in the ores, but the following minerals have been identified: malachite, azurite, anglesite, cerussite, smithsonite, bismuthite, cerargyrite, stromeyerite, chalcocite and covellite. The deposits are classified as those formed at intermediate depths and suggest a genetic relationship with an intrusion of granite porphyry. An unusual vein in the district carries, in order of deposition, pyrite, quartz, pyrrhotite, chalcocite, pyrite, malachite and siderite, the latter overlapping with malachite and chalcocite. ALICE W. FIFERSON.

Observations on secondary copper and silver sulfides in the Broken Hill Lode. F. L. STILLWELL. *Proc. Australasian Inst. Mining and Met.* No. 67, 187-210 (1927), No. 77, 71-9 (1930). The primary ore (apart from gang) consists essentially of sphalerite and galena with minor amounts of arsenopyrite accompanied by a Ni-Co-Sb mineral, chalcocite, pyrrhotite and cubanite, chalcocite and tetrahedrite, meneghinite, dyscrasite, pyrrhotite and berthierite. The oxidized zone consists of a manganese-rich ironstone capping covering large masses of cerussite frequently rich in oxidized Ag and Cu minerals. Zn salts from the primary sulfide ore have been largely removed in solution. The abundance of cerussite is partly due to the abundance in the meteoric H_2O of carbonates derived from the primary calcic ore. Only minor amounts of anglesite resulted. The surface waters contained small quantities of phosphate, derived from apatite which converted small quantities of the sulfate or carbonate into pyromorphite or Pb phosphate. At varying depths the sulfides appear. The secondary sulfides including Ag and Cu occur as coatings resembling soot on the primary. The secondary sulfides have been partially oxidized as denudation has progressed. A. H. E.

The Engels copper deposits, California. CLARENCE N. FENNER. *Econ Geol.* 25, 420-5 (1930).—The article by Knopf and Anderson (*C. A.* 25, 205) is discussed. The authors' statements in favor of gaseous transfer of mineral matter are approved, and their discussion is supplemented. ALICE W. FIFERSON.

The ores of the northern Rhodesia copper belt. ALAN M. BATEMAN. *Econ Geol.* 25, 365-415 (1930).—This belt will constitute the greatest Cu mining center of the world. Production and important mines are named. The region is a peneplain 4000 ft high. The rocks consist of an old pre-mineral basement complex unconformably overlaid by ore-bearing Roan series probably of pre-Cambrian age, intruded by younger granites and basic intrusives, the granites being the probable source of the Cu. The ore deposits consist of beds of the Roan series uniformly metallized with specks of CuS . Ores are largely sulfide and oxidation extends hundreds of ft below the water level. Chalcocite is the most important sulfide, followed by bornite and chalcocite. Linnaite is not uncommon and pyrite is rare. Eight types of chalcocite intergrowths are described and discussed, the conclusion being that the chalcocite is in part hypogene and in part supergene. Sulfides have been introduced after consolidation and folding of the rocks and have clearly replaced the rock silicates. Localization of the ore minerals in the sediments was probably due to the superior permeability of the ore beds. This Cu belt constitutes a great metallogenetic province of Cu mineralization. It is linked with the adjacent Katanga Cu region by similar features of geology and ores and the common occurrence of the rare mineral linnaite. A. W. F.

The correlation of the ore-bearing sediments of the Katanga and Rhodesian copper

belt. ANTON GRAY *Econ Geol* 25, 783-804 (1930), cf preceding abstract.—The principal Cu deposits of the Katanga and Rhodesia occur in the lower series of the System of the Katanga, known locally by various names. In view of the fact that these rocks were first studied in the Katanga, where the ore-bearing beds were called the *Série des Mines* this name has been applied by G. to the rocks over the entire Cu belt. The sulfide ore bodies of Rhodesia are found near the base of the *Série des Mines*. They lie stratigraphically below the oxidized ores of the Katanga, which are found in the dolomites of the upper Roan group, and to a lesser extent in the M'Washu. No evidence of unconformity has been found between the *Série des Mines* and the Kundelungu series and it is considered that these beds should be grouped together under the name of the Katanga system, as proposed by Van Doornick. A geologic map is included. ALICE W. IFFERSON

Some remarks on the metallogenesis of the copper beds of Katanga and northern Rhodesia. ANDRÉ DEMAY *Rev. miner.* 4, 39-46 (1930), cf preceding abstract.—The upper part of the Cu bearing beds of Katanga are chiefly replacement deposits, impregnation predominates in the lower part. Most beds are highly mineralized but spotty. In Northern Rhodesia impregnation plays the principal part in the enrichment of the sediments which are less highly mineralized but more uniform. Although the ores now being mined in Katanga are almost exclusively carbonates or oxides, the primary mineralization was undoubtedly in the form of sulfides. In Rhodesian beds, below the superficial zone of oxidation, the principal zone contains Cu sulfides. In Katanga, Fe, Mn, Co, especially Pb, Zn and even U and Ra, mined in a particular bed, together with Cu, constitute an indisputable metallogenic association. In northern Rhodesia the Cu veins are related to an intrusion of granite that penetrates the lower beds of the Roan series. In Katanga there is no direct proof of an igneous vein having affected the beds, although there are some indications that the Cu veins may have been influenced by acid magma. C. W. OWINGS

The Sherritt-Gordon copper-zinc deposit, northern Manitoba. E. L. BRUCE *Econ Geol* 25, 868-70 (1930). He discusses the communication by J. T. Wright, (*C. A.* 25, 1180) regarding H's paper (*C. A.* 23, 4910). He explains some apparent disagreements in the 2 papers, gives some additional data and states that in the main his statements and those of Wright agree. ALICE W. IFFERSON

Spectrum analysis of Mansfeld copper shale. A. CIESARZ *Chem. Erde* 5, 48-75 (1930).—To the long list of elements previously detected by ordinary analytical methods the following are now added: Sn, W, Pt, Ir, Pd and Yt, bringing the total up to 42. B. C. A.

Lead-zinc and pyrites ores of the Deutsch-Bleischarley mine, Upper Silesia. H. SCHNEIDERHÖHN *Chem. Erde* 5, 385-95 (1930).—These ores are known to contain As, they were examined by the metallographic method for As minerals. In the Pb-Zn ore, shelly blende surrounds cores of galena and jordanite ($Pb_3As_2Si_2$), but in the pyrites ore, consisting of a mixt. of pyrite, marcasite and shelly blende, no As mineral could be detected. Chem. tests showed that much As is present in the pure pyrite (no doubt in isomorphous mixt.) and only traces in the marcasite and blende. B. C. A.

Microscopic and spectroscopic investigation of the platinum-bearing rocks of the Bushveld igneous complex (Transvaal). H. SCHNEIDERHÖHN *Chem. Erde* 4, 252-86 (1929).—A description is given of the rocks and associated ore minerals from the Rustenburg, Potgietersrust and Lydenburg districts together with an account of the genesis of the deposits. In the pegmatites and contact metamorphic zone sperrylite and stibio-palladinite occur, whereas in the unaltered intrusive rocks the Pt metals are entirely present in the form of platinoferrous sulfides of other metals, free Pt, Pd and Au are encountered only as secondary occurrences in the weathered rocks. Investigation of the minerals by means of the quartz spectrograph showed that the Pt metals occurred solely in the pyrrhotite, pentlandite and nickeliferous pyrite, the greatest quantity being present in the oldest minerals. Pt and Pd are present in practically equal amounts, Ir, Rh and Ru occur to the extent of 1-10% of the Pt present, while Os is present in traces only. B. C. A.

Hydrothermal oxidation and leaching experiments; their bearing on the origin of Lake Superior hematite-limonite ores. I. JOHN W. GRUNER. *Econ Geol* 25, 697-719 (1930).—G. has previously suggested that hematite ores of the Soudan formation of the Vermilion range had been leached of SiO_2 and oxidized by hot ascending waters from a large basic magma, and new expts. support the hydrothermal hypothesis and its extension to other districts of the Lake Superior region. Leaching of SiO_2 from the Fe formation is discussed, and expts. on soln. of SiO_2 in hot H_2O and on oxida-

been identified and described a triangular boxwork being most distinctive. Leached outcrops derived from bornite generally carry a "relief" limonite—a characteristic product, orange colored with soft texture on fresh surfaces. Two limonite types derived from oxidation of tetrahedrite have been identified and described. That called contour boxwork is more distinctive. Leached outcrops from tetrahedrite usually carry local incrustations of Sb oxides that assist in identification. A W F

A mother lode gold ore. CARLTON D HOLIN. *Econ Geol* 25, 348-55(1930) — II describes specimens from the Kennedy Mine, one of the largest of the Au mines operating on the Mother Lode, in Amador County, Cal. The relative degree of favorability of the several host minerals for Au as observed in these ores is chalcopyrite, pyrite, galena, apatite, carbonates, pyrite, sphalerite and quartz; the first 5 minerals are about equally favorable, and quartz is the least favorable. A W F

Geology of the Panamint silver district, California. F MAC MURPHY. *Econ Geol* 25, 303-25(1930) — The paper includes a history of mining in the district, its location and topography, a description of the rocks and their structure, and a discussion of the ore deposits. The occurrence of numerous clean-cut Ag bearing quartz veins is described. They outcrop in limestone, schist or slate, but those in limestone are more numerous and persistent. The principal and almost exclusive vein filling is milk-white, sometimes glassy, quartz of coarsely crystalline texture. Ag is almost wholly contained chemically bound in tetrahedrite, the principal sulfide mineral. Oxidation and supergene enrichment products are not abundant in the ores, but the following minerals have been identified: malachite, azurite, anglesite, cerussite, smithsonite, bismuthite, cerargyrite, stromeyerite, chalcocite and covellite. The deposits are classified as those formed at intermediate depths and suggest a genetic relationship with an intrusion of granite porphyry. An unusual vein in the district carries, in order of deposition, pyrite, quartz, pyrrhotite, chalcopyrite, marcasite and siderite, the latter overlapping with marcasite and chalcopyrite. ALICE W EFFERSON

Observations on secondary copper and silver sulfides in the Broken Hill Lode. F L STILLWELL. *Proc Australasian Inst Mining and Met* No 67, 187-219(1927), No 77, 71-9(1930) — The primary ore (apart from gang) consists essentially of sphalerite and galena with minor amts of arsenopyrite accompanied by a Ni-Co-Sb mineral chalcopyrite, pyrrhotite and cubanite, chalcocite and tetrahedrite, meneghinite, dyscrasite, pyrrargyrite and berthierite. The oxidized zone consists of a manganeseiferous ironstone capping covering large masses of cerussite frequently rich in oxidized Ag and Cu minerals. Zn salts from the primary sulfide ore have been largely removed in solution. The abundance of cerussite is partly due to the abundance in the meteoric H₂O of carbonates derived from the primary calcitic ore. Only minor amts of anglesite resulted. The surface waters contained small quantities of phosphate, derived from apatite, which converted small quantities of the sulfate or carbonate into pyromorphite or Pb phosphate. At varying depths the sulfides appear. The secondary sulfides including Ag and Cu occur as coatings resembling soot on the primary. The secondary sulfides have been partially oxidized as denudation has progressed. A H E

The Engels copper deposits, California. CLARENCE N FENNER. *Econ Geol* 25, 420-5(1930) — The article by Knopf and Anderson (*C A* 25, 205) is discussed. The authors' statements in favor of gaseous transfer of mineral matter are approved, and their discussion is supplemented. ALICE W EFFERSON

The ores of the northern Rhodesia copper belt. ALAN M BATEMAN. *Econ Geol* 25, 365-415(1930) — This belt will constitute the greatest Cu mining center of the world. Production and important mines are named. The region is a peneplain 4000 ft high. The rocks consist of an old pre-mineral basement complex unconformably overlaid by ore containing Roan series, probably of pre-Cambrian age, intruded by younger granites and basic intrusives, the granites being the probable source of the Cu. The ore deposits consist of beds of the Roan series uniformly metallized with specks of CuS. Ores are largely sulfide and oxidation extends hundreds of ft below the water level. Chalcocite is the most important sulfide followed by bornite and chalcopyrite. Linnaeite is not uncommon and pyrite is rare. Eight types of chalcocite intergrowths are described and discussed, the conclusion being that the chalcocite is in part hypogene and in part supergene. Sulfides have been introduced after consolidation and folding of the rocks and have clearly replaced the rock silicates. Localization of the ore minerals in the sediments was probably due to the superior permeability of the ore beds. This Cu belt constitutes a great metallogenetic province of Cu mineralization. It is linked with the adjacent Katanga Cu region by similar features of geology and ores and the common occurrence of the rare mineral linnaeite. A W F

The correlation of the ore-bearing sediments of the Katanga and Rhodesian copper

belt. ANTON GRAY *Econ Geol* 25, 783-804 (1930), cf preceding abstract—The principal Cu deposits of the Katanga and Rhodesia occur in the lower series of the System of the Katanga, known locally by various names. In view of the fact that these rocks were first studied in the Katanga where the ore bearing beds were called the *Série des Mines*, this name has been applied by G. to the rocks over the entire Cu belt. The sulfide ore bodies of Rhodesia are found near the base of the *Série des Mines*. They lie stratigraphically below the oxidized ores of the Katanga, which are found in the dolomites of the upper Roan group and to a lesser extent in the M Washu. No evidence of unconformity has been found between the *Série des Mines* and the Kundelungu series and it is considered that these beds should be grouped together under the name of the Katanga system, as proposed by Van Doorninck. A geologic map is included. ALICE W. IFFERSON

Some remarks on the metallogenesis of the copper beds of Katanga and northern Rhodesia. ANDRÉ DEMAY *Rev. univ. mines* 4, 38-46 (1930), cf preceding abstract—The upper part of the Cu bearing beds of Katanga are chiefly replacement deposits, impregnation predominates in the lower part. Most beds are highly mineralized but mostly in Northern Rhodesia impregnation plays the principal part in the enrichment of the sediments which are less highly mineralized but more uniform. Although the ores now being mined in Katanga are almost exclusively carbonates or oxides, the primary mineralization was undoubtedly in the form of sulfides. In Rhodesian beds, below the superficial zone of oxidation, the principal zone contains Cu sulfides. In Katanga, Fe, Mn, Co, especially Pb, Zn and even U and Ra, mined in a particular bed, together with Cu, constitute an indisputable metallogenic association. In northern Rhodesia the Cu veins are related to an intrusion of granite that penetrates the lower beds of the Roan series. In Katanga there is no direct proof of an igneous vein having affected the beds, although there are some indications that the Cu veins may have been influenced by acid magma. C. W. OWINGS

The Sherritt-Gordon copper-zinc deposit, northern Manitoba. E. L. BAUCE *Econ Geol* 25, 848-70 (1930). B. discusses the communication by J. F. Wright, (*C. A.* 23, 1186) regarding B's paper (*C. A.* 23, 4910). B. explains some apparent disagreements in the 2 papers, gives some additional data and states that in the main his statements and those of Wright agree. ALICE W. IFFERSON

Spectrum analysis of Mansfeld copper shale. A. CIESARZ *Chem. Erde* 5, 48-75 (1930).—To the long list of elements previously detected by ordinary analytical methods the following are now added: Sn, W, Pt, Ir, Pd and Yt, bringing the total up to 42. B. C. A.

Lead-zinc and pyrites ores of the Deutsch-Bleischarley mine, Upper Silesia. H. SCHNEIDERHÖHN *Chem. Erde* 5, 385-97 (1930).—These ores are known to contain As, they were examined by the metallographic method for As minerals. In the Pb-Zn ore, shelly blende surrounds cores of galena and jordanite ($Pb_3As_2Si_2$), but in the pyrites ore, consisting of a mixt. of pyrite, marcasite and shelly blende, no As mineral could be detected. Chem. tests showed that much As is present in the pure pyrite (no doubt in isomorphous mixt.) and only traces in the marcasite and blende. B. C. A.

Microscopic and spectroscopic investigation of the platinum-bearing rocks of the Bushveld igneous complex (Transvaal). H. SCHNEIDERHÖHN *Chem. Erde* 4, 252-86 (1929).—A description is given of the rocks and associated ore minerals from the Rustenburg, Potgietersrust and Lydenburg districts, together with an account of the genesis of the deposits. In the pegmatites and contact metamorphic zone sperrylite and stibio-palladinite occur, whereas in the unaltered intrusive rocks the Pt metals are entirely present in the form of platinumiferous sulfides of other metals; free Pt, Pd and Au are encountered only as secondary occurrences in the weathered rocks. Investigation of the minerals by means of the quartz spectrograph showed that the Pt metals occurred solely in the pyrrhotite, pentlandite and nickeliferous pyrite, the greatest quantity being present in the oldest minerals. Pt and Pd are present in practically equal amounts. Ir, Rh and Ru occur to the extent of 1-10% of the Pt present, while Os is present in traces only. B. C. A.

Hydrothermal oxidation and leaching experiments; their bearing on the origin of Lake Superior hematite-limonite ores. I. JOHN W. GRUNER *Econ Geol* 25, 697-719 (1930).—G. has previously suggested that hematite ores of the Soudan formation of the Vermillion range had been leached of SiO_2 and oxidized by hot ascending waters from a large basic magma, and now experiments support the hydrothermal hypothesis and its extension to other districts of the Lake Superior region. Leaching of SiO_2 from the Fe formation is discussed, and experiments on soln. of SiO_2 in hot H_2O and on ox-

dation of minerals containing Fe^{++} are described. The SiO_2 was found to be very sol., and tables are given showing the solubilities of 4 ferrous minerals, SiO_2 from gal. bro being found less sol. than quartz. It was found that the oxidation of ferrous minerals by water at high temp. begins at or below 200° . The expts. and methods used are described in detail. *Il. Ibid* 837-67.—Expts. show that SiO_2 from SiO_2 gal., chalcedony, and greenalite tacomite is sol. to an extraordinary extent in water at temps. from 200 to 300° . The soly., except of SiO_2 from gal. bro is many times as great at 200° as at room temp. Oxidation of Fe^{++} in minerals takes place in hot H_2O or steam in the absence of air at temps. as low as 160° provided the H given off in the decompn. can escape. Siderite is oxidized to magnetite, and magnetite to hematite. Expts. support the hypothesis that Lake Superior hematite limonite ores were oxidized and leached hydrothermally. These ores have similar counterparts in the leached ores of South Africa. Arguments supporting the weathering hypothesis and G's hypothesis are discussed and the latter is upheld. ALICE W. EPPERSON

The iron deposits of the Sierra de Imataca, Venezuela. ERNEST F. BURCHARD. *Econ. Geol.* 25, 549-561 (1930).—A paper by Zubizarri on Pao Fe ore deposits is discussed in connection with one by B. (*A. I. M. M. E. Tech. Pub.* 275, Jan., 1930). Similarities and dissimilarities in the ores of the 2 regions are pointed out. A. W. E.

Graphic intergrowth of niccolite and chalcopyrite, Worthington Mine, Sudbury, CARL LAUSEN. *Econ. Geol.* 25, 376-64 (1930).—The general geology of the region and the ore deposits are described, and the origin of the graphic texture and sequence in the deposition of the ores is discussed. The mineralization is described as taking place in 3 stages: the 1st was a high temp. hydrothermal stage, producing magnetite, pyrrhotite, pentlandite, chalcopyrite, the 2nd and 3rd were moderate temp. hydrothermal stages, producing resp. *graptolite*, mancherite, niccolite-chalcopyrite intergrowths and pyrite, sphalerite and galena. ALICE W. EPPERSON

A qualitative and quantitative determination of the ores of Cobalt, Ont. L. ELLIS THOMSON. *Econ. Geol.* 25, 470-505 (1930).—Methods of analysis are given, and development of new etching reagents is described. Geology of the region is discussed, and detailed descriptions are given of specimens from Mining Corporation, Nipissing and O'Brien Mines. The old mineralogical title of smaltite-chloanthite for the hard white mineral associates of the native Ag and niccolite of this region is no longer applicable, these 2 minerals being exceeded in rarity only by the mineral *graptolite*. Cobaltite is much the commonest of these minerals. Common structures are noted for the 9 minerals, listed in order of prevalence as follows: cobaltite, arsenopyrite, saffronite, skutterudite, rammelsbergite, kottunite, smaltite, chloanthite, *graptolite*. *Il. Ibid* 627-32.—Eight new microchemical reagents have been devised for differentiation of the 9 hard white Ag minerals which are common in this district. Crystal habits, modes of occurrence, interrelationships and associations with Ag of each of these minerals are discussed. Relative abundance of the common minerals is given, smaltite and chloanthite are placed among the rarer minerals of the group with cobaltite at the top of the list. Relation between quant. proportions of the other common minerals of this district and their proximity to the diabase. *Il.* worked out and compared to a similar relation for native Ag. It appears that larger quantities of mixed arsenopyrite, cobaltite, kottunite, rammelsbergite or skutterudite have been favorable to the deposition of native Ag, while niccolite, tetrantite and saffronite in larger quantities are inimical to it. The occurrence of rarer minerals is recorded as observed by other investigators. Details of ore groups, as outlined by previous observers, are amplified to include the occurrence of such minerals as saffronite and marcassite, which are new for the district, and to fit the commoner occurrence of other minerals which were heretofore considered rare. ALICE W. EPPERSON

A qualitative and quantitative determination of the ores of Cobalt, Ont. M. N. SMYTH. *Econ. Geol.* 25, 774-1' (1930).—cf. preceding abstract.—The work of Thomson is discussed at length, and S. makes suggestions giving his own methods of detn. of the hard arsenides. ALICE W. EPPERSON

Origin of chromite deposits. F. E. KREEP. *Econ. Geol.* 25, 425-6 (1930), cf. C. A. 25, 899.—The paper of Fisher (*C. A.* 24, 1219) is discussed. K. agrees with F. as to the possible cause of the chromite in the Great Dyke, but disagrees with the statement that the Great Dyke is younger than the rocks of the Bushveld Complex and gives evidence that the 2 intrusives were more or less contemporaneous. A. W. E.

The chrome deposits of Unvukwes district. F. E. KREEP. *S. African Mining and Eng. J.* 41, Pt. II, 165-6 (1930).—The chrome ore occurs in seams in serpentinized enstatite-olivine rock and in websterite. The ore was formed by magmatic segregation.

The seams are expected to extend to some depth without change. The uses of chrome ore are discussed. ALICE W. LIPPERSON

Supergene cassiterite in tin veins. FRIEDRICH AHPFELD. *Leon Geol* 25, 546-8 (1930) —The paper by Singenwald (C. A. 25, 899) is discussed. Results of investigations in Bolivia are given. Meteoric waters are not capable of dissolving cassiterite since they (in Bolivia) are always acid. All surface waters are not present in Bolivia since they (in Bolivia) are always acid. ALICE W. LIPPERSON

Supergene cassiterite in tin veins. J. B. SCRIVNER. *Leon Geol* 25, 663-4 (1930), cf. preceding abstract —The citation of bony tissue replaced by cassiterite is thought to refer to certain Cornish stag horns which were said to be impregnated with SnO_2 , but which were found by the writer to contain no Sn, but to have grains of quartz in cavities of the bone, evidently washed in by water. No evidence was found by the writer in Cornwall, or in Malaya, of solns of cassiterite by ground water. ALICE W. LIPPERSON

Origin of the strontium in the strata of the lower Muschelkalk and Rot formations near Jena. K. DINGER. *Chem Erde* 4, 167-77 (1929) —The celestite occurring in certain strata of the lower Muschelkalk and Rot formations is not primary, but is due to secondary enrichment. The Sr has been transported in soln., probably as $\text{Sr}(\text{HCO}_3)_2$, and pptd. as sulfate by interaction with gypsum since the deposits of celestite occur only where this mineral is or has been, present. Spectroscopic and chem. examn. of the various strata of the above 2 formations showed the presence of traces of Sr in practically all cases, derived initially from sea water through the agency of certain radiolaria which abstract it to form their shells. B. C. A.

Search for phosphate deposits in U. S. S. R. A. D. ARKHANGELSKII. *Compt rend acad sci U. S. S. R.* 1930A, 97-101 —Possible sources of phosphates in Russia have not been so far thoroughly investigated. A thorough study is recommended, especially in Crimea, Caucasus and Turkestan, not only of known types of deposits, but also of general distribution of phosphates in sedimentary rocks. Most of the deposits known at present are of low P content and are scattered among other minerals over a large area in thin layers. Formation of phosphates from limestones by means of reaction with NH_3 , phosphate originating from deposits of guano is outlined, and possibility of this phenomenon in Russia given consideration. ROBERT SCHLASS

Asbestos mulling in Rhodesia. F. L. KERR. *S. African Mining and Eng. J.* 41, Pt. II, 89-90 (1930), cf. C. A. 24, 3198 —Serpentinization and the formation of chrysotile asbestos in the Great Dyke of Rhodesia are probably due to surface waters. The zones are localized in zones of weakness. Of the rock passing the grizzly at the mill head, 8-12% passes to the cleaning mill as fiber and grit. ALICE W. LIPPERSON

Genesis of the emery deposits near Peekskill, New York. JOSEPH L. GILLSON AND JOSEPH L. A. KANIA. *Econ Geol* 25, 506-27 (1930) —These deposits, still mined on a small scale, are located 35 mi. N. of N. Y. City on the east bank of the Hudson, along and on both sides of the contact of the Manhattan schist and Cortlandt intrusives G and K, present data disproving the assimilation theory of the deposits and proving that they are contact metamorphic in origin and were formed by gaseous or liquid emanations from the magma reservoir which passed upward through the already solid border of the igneous mass and into the schist. A detailed mineralogical and petrographical description of the rocks and ores is given. ALICE W. LIPPERSON

The origin of the talc and soapstone deposits of Virginia. J. D. BURFOOT, JR. *Econ Geol* 25, 805-26 (1930) —The geographic distribution of these deposits is discussed, and their characteristics described. The presence or absence of talc, soapstone or steatite in this region is thought to depend, in part at least, on the occurrence of the correct segregation product—pyroxenites, peridotites or dumites, on the presence of feeding channels leading to these rocks, and on the invasion of the correct type of solns., that is, solns. of the correct chem. character with the requisite temp. and pressure. The condition of differential stress is not necessary for the formation of talc. The mineral association and paragenesis seem to indicate that talc is formed under deep or intermediate vein zone conditions. ALICE W. LIPPERSON

Origin of the white clays of Tuscaloosa age (Upper Cretaceous) in Alabama, Georgia and South Carolina. GEORGE I. ADAMS. *Econ Geol* 25, 621-6 (1930) —A. believes that the white clays in the district were not the result of marine deposition, as held by Newmann. Deposition of clayey material in small shallow fresh-water lakes is more commensurate with the extent of the white clay bodies, but does not explain the massive character of the beds. It is suggested that the massive character and purity of the clay are the result of the local alteration of deposits of impure clays. No attempt is made to explain the chem. processes which are involved. A satisfac-

tory explanation must be applicable to all the white clays in the region which are similar in character, whether their sources were the crystallines and metamorphics, or the sedimentary areas, and whether they occur in the Tuscaloosa or other formations

ALICE W. EPPERSON

Clays of the Jackson Purchase region, Kentucky. JOSEPH K. ROBERTS. *Econ. Geol.* 25, 832-6(1930)—The geological formations of this region range from the limestone and chert of Mississippian to the Recent surficial deposits. Detailed descriptions are given of a number of the clays. Tables of chem. analyses are included

ALICE W. EPPERSON

Geology of Tshasoff-Jar fireclay deposits. P. A. ZEMYATCHEVSKII. *Trans. Ceramic Research Inst. (Moscow)* 1929, No. 20, 7-29 (in German 182-3)—The fireclay occurs in beds in Tertiary Pollawa sands. Intact, horizontal layers are found, frequently curved or folded, sometimes compressed or even torn off the main mass. Perpendicular sections are not uniform in sand content, color and Fe(OH)₃ inclusions. There is also included lowgite and alunitite, usually in white and yellow nodules from the size of a pea to 20-30 cm. The clay bed is 4.5-6 m thick, sandy at the top. The area worked is about 1200 hectares, by several concerns and at 24 places. Without careful sorting, the quality of the clay suffers from the presence of lowgite, which on burning gives SO₂ and SO₃, causing swelling and blistering. Na₂SO₄ remaining from the lowgite is a good flux, and if coned at certain points can cause holes to form in the ceramic object. It is proposed to levigate the clay to make it uniform and to purify it. Based on surveys, about 15 million tons of clay are available. E. M. SYMONES

Composition of clay from the Tshasoff-Jar deposits. V. ISEYUL. *Trans. Ceramic Research Inst. (Moscow)* 1929, No. 20, 30-37 (in German 184-7), cf. preceding abstr.—Examined by microscope and detn. of *ns* showed the clay to be mainly kaolinite in the form of shredded flakes, muscovite in colorless scales and slightly rounded quartz grains. Tourmaline and limonite were present, the former in free, prismatic particles or embedded in quartz grains, the latter as compact clusters or granules distributed throughout the clay. Rutile grains were found in only the finest clay fractions, zircon in long grains or whole crystals usually in the coarse fraction. Strongly weathered feldspar, lowgite and alunitite were found in rounded grains, and biotite in brown plates. The clay was sepd. by the Sabarun and Robinson method into fractions > 0.05 mm., 0.05-0.01 mm. and < 0.01 mm. Chem. analyses showed high alkali content. About 1-2% TiO₂ was always found, but not much Fe₂O₃, CaO and MgO. SO₃ was present, derived from lowgite and alunitite. The org. content corresponded to humus. The crude clay reacts very little with Na₂CO₃ soln. or with 2 and 10% HCl, but ignited clay reacts easily, Fe₂O₃, CaO, MgO and alkali being extd. The clay is a kaolinite mica-quartz rock. E. M. SYMONES

Mineralogical composition of Doubrovka kaolin and its fractions by rational analysis. T. KRASENCEVA. *Trans. Ceramic Research Inst. (Moscow)* 1929, No. 21, 3-25 (in German 26-7)—The change of mineral compn. of kaolin on alteration from coarse to fine fractions was detd. and mineral impurities were identified. The kaolin was decompd. by 10% HCl after dehydration at 600° followed by a 2-hr. extn. by 5% Na₂CO₃ soln. Microscopic examn. of kaolin showed the presence of kaolinite, quartz, mica and feldspar. Ease of decompn. of mica and feldspar is not given in the literature. Their proportions were estd. by microscope count. Mica was 95% decompd. by the HCl and Na₂CO₃, and feldspar about 23%. Tables show the mineralogical compn. of unground kaolin and its coarser fractions. In the finer fractions kaolinite rises, but mica, feldspar and quartz diminish. E. M. SYMONES

Bohemian clay. G. LENCE AND E. KÖHLER. *Chem. Erde* 4, 459-60(1930).—A clay from the Tertiary lignite beds was analyzed in bulk and also the portion sol. in HCl after ignition at 600°. The dehydration curve is plotted. It contains kaolin 47.58 with amorphous material 18.94%, quartz 40% and some mica, pyrite and org. matter. B. C. A.

A new locality of gaylussite in eastern Mongolia with associated natural soda. K. NINOMY. *Econ. Geol.* 25, 758-63(1930)—The topography of the region is described, and the occurrence of Na minerals and of gaylussite is given, with typical analyses. The origin of these minerals is discussed, but no definite theory is formulated. ALICE W. EPPERSON

Geological aspects of the formation of coal. CYRIL S. FOX. *Proc. 16th Indian Science Congress, Geology Sect., October 25, 1929; Fuel in Science & Practice* 9, 548-50(1930)—Stratigraphical and petrological data obtained from field investigations and lab. studies of Indian coals are presented in a detailed discussion of their occurrence, quality, origin, constitution and formation. From these data and the results of nu

merous other researches—39 references are included—It draws 16 conclusions concerning the formation of coal. The main processes in the conversion of peat into primary coal are considered to be of a chem character and involve all the plant components. Coalification proceeded under water, under anaerobic conditions, sometimes before deposition of overlying sediments. Primary coal became heated because of pressure and was thus changed into bituminous or anthracite coal depending upon the degree of carbonization. In some cases where carbonization proceeded far enough locally, as a result of catalytic action in capillary spaces fusain was formed. D. A. R.

Coal as a recorder of incipient rock metamorphism. MARCUS R. CAMPBELL. *Econ Geol* 25, 675-96(1930)—Evidence is given to indicate that 6 stages in the metamorphism of coal can be recognized and explained by the processes that have been active in the earth coincident with the development of plant life on the land, and that at least 6 of these stages have left few if any marks on the compn or phys appearance of the sediments associated with the coal. They may, therefore, be classed as incipient stages of rock metamorphism. ALICE W. EPPERSON

Studies in the development of Dakota lignite. V. Extraction and study of the benzene-soluble portion of Dakota lignite. I. I. HARRIS, C. F. BELCHER AND A. W. GAUGER. *Ind Eng Chem* 23, 199-201(1931). Lignite from the Velva deposit in N. D. was extd by the Licher (C. I. 11, 1739) and Bone (C. A. 23, 1931) methods and a study made of the CCl₄-sol portion. Up to 43% of the dry, ashless coal substance was extd by the Bone method. That portion of the ext obtained up to 17 atm has a compn closely analogous to Morkin wax from German brown coal. A. H. E.

Some chemical aspects of the origin of petroleum. S. C. LIND. *Science* 73, 19-22(1931)—The chem complexity of petroleum is little less miraculous than its abundance. This discussion is directed toward the origin of the complexity rather than the initial source of natural hydrocarbons. Processes, either thermal or ionic, are now known by which progression both up and down the hydrocarbon series is effected. This leads directly to the complexity found in natural petroleum, and also found in synthesized ones. Consequently, the starting material, whether of vegetable, animal or mineral source, does not need to be a complex mixt, but may be a single chem species, from which a high degree of complexity is obtained by steps which appear simple when chem and thermodynamic properties of hydrocarbons are considered. The simplicity of such a mechanism may lend indirect support to the old idea of inorg origin from one or a few hydrocarbon gases such as might be produced by action of water on metallic carbides in the earth's interior. On the other hand, it does not preclude animal or vegetable origin, but strongly suggests that the primary material, whether gaseous, liquid or solid, is later subjected to thermal (or ionic) agents or both, which produce the complexity found in nature. ALICE W. EPPERSON

Study of the black shale overlying the cap rock of the Cromwell sand in relation to the origin of the Cromwell oil dome, Oklahoma. OLIVER R. GRAVE. *Econ Geol* 25, 326-47(1930)—An investigation of 54 samples of this shale has shown that its lithologic properties and its compn are as const as can be expected of any rock. No variations occur which can be correlated with local metamorphism produced by structure forming forces. Slight differences in color are independent of the structure and seem to accompany changes in compn (especially in carbonate content) which may be best explained on the basis of sedimentation. Although shales from different stratigraphic horizons show increase in sp gr with depth, sp gr of the shale from a single horizon is constant, regardless of structure. Gardner's suggestion that a careful study of properties and compn of the shale in the low dome-like structures of the Mid Continent field may prove the presence of slight localized metamorphism has not been substantiated, and it is unlikely that local forces of the type postulated by Gardner could have produced the Cromwell dome. ALICE W. EPPERSON

Limestone oil reservoirs of the northeastern United States, and of Ontario, Canada. A. N. MURRAY. *Econ Geol* 25, 452-69(1930)—On theoretical grounds the most effective, if not the only, way in which sufficient porosity can be developed in limestone to provide a reservoir of commercial importance, is as a result of soln above a water table. The porous limestone reservoirs in the northeastern U. S. and Ont. all have been shown to have been subjected to erosion before the time of deposition of the formation now overlying them, with the possible exception of the limestones of Traverse age in Mich. It is therefore suggested that erosion caused the development of porosity in the case of all these limestone reservoirs. ALICE W. EPPERSON

Some properties of limestone as a reservoir rock. W. V. HOWARD AND W. W. LOVE. *Econ Geol* 25, 720-36(1930)—The expts described indicate that oil will migrate through limestone, but that most of the oil is absorbed by the rock, particularly

the lighter fractions of oil. In a thick reservoir heavy residues may collect in tighter zones within the reservoir with the result that porous limestone is divided into "pays" separated by an asphaltic limestone, as well as by other impervious beds. Such division into zones makes difficult the application of expl. results to field conditions. It appears that normal production methods will remove from the area tributary to any well all oil which that well can be expected to produce, regardless of rate of production. The low sp. gr. typical of oils from limestone reservoirs appears to be the result of absorption of lighter fractions by the rock.

Chemical and petrological investigation of bituminous rocks of various ages from northern Germany. F. HRENTHAL. *Chem. Erde* 4, 343-68 (1929).—Bituminous material from 4 different formations (Alluvial, Diluvial, Tertiary and Cretaceous) has been examd. chemically and petrologically. The bitumens from these marine sediments show increasing C content with age, but unlike those of terrestrial origin there is no corresponding decrease in the H content. Elementary S was found in the specimen from the Alluvium and a stable "metabiten" was detected in the Tertiary rock. B. C. A.

A tholeiitic phase of the quartz-dolerite magma of central Scotland. FREDERICK WALKER. *Mineralog. Mag.* 22, 368-76 (1930).—The rocks near Dalmeny and Kinkell are described, with chem. analyses, and shown to contain chlorophacite. The analyses and ws. indicate that the residual glass is of acid compn. in both cases. B. C. A.

Granitic rocks of Tsukuba district and their associated injection rocks. KEN-ICHI SUGI. *Japan J. Geol.* 8, Nos. 1/2, 29-112 (1930).—Around the granite rocks of the district hornfelses and injection rocks are extensively developed. Biotite granites (granodioritic) constitute the main body of the igneous rocks, being assoc. with aplitic granites, aplites and pegmatites which occur as small sheets and are especially abundant near the contact borders of the granites. These granitic rocks may have intruded during the Mesozoic era. The metamorphic rocks are believed to have been derived from the sedimentary complex of the Chichibu series, the metamorphism being due to the igneous intrusions. Of the hornfelses which occupy the eastern part of this area, the predominant types are those derived from argillaceous sediments andalusite- or cordierite hornfels, which show gradual transitions to the injection rocks of the western area. The latter, derived largely from argillaceous sediments, are grouped in 4 main types, viz. sillimanite-biotite, biotite- with shimmer aggregates cordierite-biotite and biotite fels. There are also metamorphic rocks derived from Ca-Mg bearing sediments, which are almost free from igneous injection but not from its effects. A few chem. analyses are given. W. L. HILL.

Geological and petrographical examination of the granites quarried in Silesia. W. BECKER and T. MACHR. *Chem. Erde* 5, 412-36 (1930).—Details are given of the stone from several quarries in the neighborhood of Strehlen, including the jointing of the rock, the size of grain of each constituent mineral and the relative proportions of each mineral. These are points of importance in the quarrying of the stone and in its suitability for various purposes. B. C. A.

Phonolites and trachytes of the Laacher See district. W. AHRENS. *Chem. Erde* 5, 1-21 (1930).—Trachytoid phonolite tuffs at Burgbrohl, belonging to the Alluvial period, are of later age than the main mass of trachytes of the Laacher See district. The contained blocks of trachytoid phonolite lava are described petrographically with chem. analyses. The analyses are compared with published analyses of the trachytes and plotted on diagrams showing that at this later period the magma became more acid. B. C. A.

Mineralogy and petrography of the Rhön Mountains. F. HEIDE. *Chem. Erde* 5, 396-411 (1930).—Petrographical descriptions are given of phonolite rocks, with one chem. analysis. B. C. A.

Basalt from the Atlantic Ocean. C. W. CORRENS. *Chem. Erde* 5, 76-86 (1930).—A description with chem. analysis is given of a fragment (62.7 g.) of basalt raised from the floor of the Atlantic at a depth of 2000 m. in lat. 1°56' S., long. 12°40' W. It has a glassy crust which has been acted on by sea water. B. C. A.

Basalt of the Stöfel, Westerwald [Nassau]. E. LEHMANN. *Chem. Erde* 5, 319-72 (1930).—From this hill of Miocene basalt 600,000 tons of stone are quarried annually. The basalt is penetrated by dikes of essentite-porphry, which is a product of differentiation of the basaltic magma. Detailed petrographical descriptions with chem. analyses are given of these rocks. B. C. A.

Dolomitization of Middle Devonian [limestone] in the Eifel. P. MÜLLER. *Chem. Erde* 4, 431-54 (1930).—Many analyses are given of the dolomite rock in the Gerolstein

tain. They show in the majority of cases 50-96% of dolomite ($\text{CaCO}_3 \cdot \text{MgCO}_3$), but in a few streaks only 1-2%. The rock is a coral reef formation; the dolomitization took place contemporaneously by the action of the sea water. The unfoliated streaks represent lagoon deposits. B C A

Rocks of Gomera, Canary Islands. W. MCLURE. *Chem. Erde* 4, 369-61 (1930).—The Tertiary or post Tertiary volcanic rocks of the island of Gomera are underlain by a series of extensively weathered rocks which were believed by Gage (C. A. 20, 311 f.) to represent a much older "Grundgebirge." A detailed examination with several analyses of these rocks shows that they are only the weathered equivalents of the ordinary types of volcanic rocks which include basalts, trachytes and trachytic phonolites. Analyses are also given of the portions of the weathered rocks sol in HCl. B C A

Red marls of the Trias formation. H. P. KRAUSS. *Chem. Erde* 4, 188-207 (1930). A number of marls from the red Trias formation have been subjected to chemical and rational analysis, and their mineral composition has been noted. After subtraction of elastic minerals (quartz, feldspar and carbonates) and freedom of the results to 100% it is found that the ratio of Al_2O_3 to SiO_2 in the residual material is practically the same as that in specimens of terra rossa. It is therefore concluded that both these substances represent similar products of weathering. B C A

"Terra rossa" as residue from solution of marine limestones. W. H. GRAY JR. LEWIS. *Chem. Erde* 4, 174-87 (1930). A series of comparative analyses of "terra rossa" from a number of localities and of material left after the solution of various limestones is adduced as additional evidence in support of Lewis' view that the first named substance is to be regarded as a residual residue of certain marine limestones. B C A

"Terra rossa" as solution residue of marine limestones. J. BLANCH. *Chem. Erde* 5, 43-7 (1930). The analyses by Lewis (preceding abstract) support the view first expressed by B. in 1916. B C A

Natural weathering and a comparison of chemical and natural weathering of building stones. F. KATZ. *Chem. Erde* 4, 290-312 (1930). In a study of the effects of natural weathering in a region, the properties of the underlying strata are important, since the nature of these often greatly influences the effects due to climatic weathering, certain cases furnish instances of this kind. The weathering of building stones caused by various gases present in the atmosphere is fully discussed, and the analogy with natural weathering is traced. Analysis of a block of sandstone taken from Regensburg Cathedral showed that of the salts formed by weathering, CaSO_4 tended to be concentrated in the outermost layers, while MgSO_4 and Mg chlorides had penetrated deeply into the stone. The presence of alumina and silica in the soil salts due to weathering was also noted, but it is uncertain in what form they occur. Chemical weathering is intensified by light, and building stones exposed to full light are more rapidly corroded than those in a more shaded position. B C A

Bleaching processes [in rocks]. H. HARRASSOWITZ. *Chem. Erde* 5, 146-61 (1930).—Many different geological processes may give rise to much the same result. Exogenic leaching and endogenic leaching are distinguished, the former is due to ordinary weathering at the surface, and the latter to fumarole and hydrothermal actions. By the removal of certain constituents the rocks may be altered to kaolin, bauxite, alunite, etc., or even to a pure quartz sand. B C A

Weathering of shell limestone and soil formation near Jena. W. HOFER. *Chem. Erde* 5, 165-81 (1930).—Several chemical analyses are given of limestones from different strata in the "Muschelkalk," of the weathered rocks and of the overlying soils. These show a gradual removal of lime and an accumulation of the residual constituents of the limestone. B C A

The abundance of elements of the vanadium group in eruptive rocks. G. V. HEDBERG, B. ALEXANDER AND K. WÖRSTLIN. *Z. anorg. allgem. Chem.* 194, 316-25 (1930).—Two hundred and eighty-two samples of eruptive rocks were examined for V, Cr, Ta and protoactinium content. The mean Cr and Ta content was 3.2×10^{-1} or 2.4×10^{-1} g. per gram of rock. The relation of the abundance of Cr and Ta is 2.6. The V content was 2×10^{-4} and the protoactinium content (calculated from the Ra content of 1.2×10^{-13}) 7×10^{-13} per gram. A comparison was made of the abundance of these elements of the V-group in the eruptive rocks and in meteorites. M. M.

The significance of silicate synthesis for geochemistry. WILHELM FRIEDEL. *Naturwissenschaften* 18, 1019-26 (1930).—A review of recent work in silicate chemistry. B. J. C. VAN DER HORVEN

The age of the earth. OTTO HAHN. *Naturwissenschaften* 18, 1013-6 (1930).—A review. B. J. C. VAN DER HORVEN

A study of the pedogenic processes in an area of lower paleozoic shales (JONES) 15.
Geochemical distribution laws and cosmic frequencies of the elements (GOLDSCHMIDT) 2.

STADNICKOY, GEORG Die Entstehung von Kohle und Erdöl. Die Umwandlung, organ. Substanz im Laufe geolog. Zeiträume. Hft 5/6 of "Schriften auf d. Gebiet d. Brennstoff Geologie." Stuttgart. F. Enke 254 pp. M 20

Standards and Specifications for Nonmetallic Minerals and Their Products. Washington, D. C. Supt. of Documents, Government Printing Office 683 pp. \$2.75

9—METALLURGY AND METALLOGRAPHY

D. J. DEMAREST, R. W. GILLET AND RICHARD RIMBACH

Metallurgy and the power industry. J. B. ROMER. *Metals and Alloys* 1, 858-70 (1930). —A review A. J. MONACK

Metallurgy in South Africa. S. W. SMITH. *Mining Mag* 43, 255-9, 256-41 (1930), 44, 24-31 (1931). —A survey of operations at various works, especially those for Au, Pb, Cu and Fe, with discussion of papers presented at the Empire Congress. A. BUTTS

Lead. G. W. THOMPSON. *Am. Metal Market* 38, No. 35, 3, 5, 10 (Feb. 20, 1931). —The great increase in consumption of Pb has been in the elec. industry where it is chiefly used for cable coverings. The problem in the future of the Pb and Cu industries will depend largely upon the demand created by the elec. industry. Properties of Pb and its alloys with Sb or Sn or both are discussed. W. H. BOYNTON

The Roan Antelope copper mine. A. G. MCGREGOR. *Mining Mag* 43, 270-6 (1930). —A description of the mine, mill and smelter with several flow sheets and diagrams. Initial production will be 50,000 tons of Cu yearly. A. BUTTS

Milling practice at Fresnillo. R. E. BYLER. *Mining Mag* 42, 137-47 (1930). —A description of practice in cyaniding an oxidized Ag ore and differential flotation of Pb-Ag-Zn sulfide ore. The cyanide is regenerated by treating the soln with SO₂ and absorbing the liberated HCN in alk. soln. A. BUTTS

Amulet flotation mill practice. W. G. HOBLER. *Can. Mining Met. Bull.* No. 226, 295-306 (1931). —Operations at the new 300-ton mill treating Cu-Zn-Fe ore by selective flotation are described, and the use of various reagents is discussed. A. BUTTS

The new smelter and concentrator of the International Nickel Company at Copper Cliff, Ontario. L. M. SPERIDIAN. *Eng. J. (Can.)* 13, 682-7 (1930). —After coarse crushing by hand at the mines, the ore is shipped to Copper Cliff where it is crushed to 200-mesh and sent through a wet concn. to remove the rock. In the concn. process the sulfides are floated, and the lighter rock material sinks to the bottom. By selective flotation a rough sepn. is made of the Cu and Ni. Smelting comprises roasting in multiple-hearth roasters, in reverberatory furnaces and blowing in basic-lined Bessemer converters. The product is an 80% Cu-Ni mat. The layout of the smelter and concentrator is shown in a diagram. W. H. BOYNTON

The Hollinger assay office. W. R. DODGE. *Can. Mining Met. Bull.* No. 225, 115-22 (1931). —The equipment and work of the lab. of the Hollinger Consolidated Gold Mines are described. A. BUTTS

Review of iron and steel literature for 1930. E. H. MCCLELLAND. *Blast Furnace Steel Plant* 19, 258-61 (1931). E. H.

Recent developments and researches in cast iron. J. W. DONALDSON. *Roy. Tech. Coll. Met. Club J. (Glasgow)*, 1929-30, No. 7, 38-40. E. H.

Evaluation and testing of the properties of raw materials and key-products in the iron industry. A. WAGNER. *Stahl u. Eisen* 50, 653-68 (1930). —The present state of our knowledge of the properties of coal, coke and Fe ore used in the production of Fe and steel and of the pig Fe produced from various types of ore is reviewed, and modern methods of detg. the value of these materials for their particular purposes are briefly discussed. Lab. tests of the coining power of coal and chem. analyses of Fe ore do not give sufficient information as to the value of the coke and ore for metallurgical purposes, as no account is taken of their phys. condition, which is of considerable importance in detg. their behavior in smelting. Similarly, chem. analysis of pig Fe affords little information as to its value for subsequent operations, but, in combination with an examn. of its micro- and macro-structure and a detn. of its mech. properties after remelting under standard conditions, the analysis gives a fairly accurate idea of the com. value of the metal. B. C. A.

Calorific value, heat and gas flow, the physical bases of metallurgical processes. H. BANSKY *Stahl u. Eisen* 50, 668-78(1930) Diagrams have been constructed to show the heat balance in various stages of the smelting of the blast furnace, gas producer, acid and basic open hearth furnaces, converter and sponge iron producer and the effect of preheating the gas and air is illustrated. A knowledge of the calorific value of the fuel and the rate of flow and heat content of the various gases is shown to be of paramount importance in the correct and efficient operation of the various processes. The resistance of the column of charge in the blast furnace to the flow of gas through the furnace acts to a large extent the output of the furnace, and any factor which increases this resistance reduces the efficiency of the furnace. One of these factors is the deposition of C from the flue gases in the upper part of the blast furnace shaft. An apparatus is described and illustrated for the determination of the magnitude of this deposition from blast-furnace gas. B. C. A.

The carbon-oxygen equilibrium in liquid iron. H. C. VACHER AND E. H. HAMILTON *Am. Inst. Mining Met. Eng. Tech. Pub. No. 409*, 11 pp (1931) The product of the C and O contents in liquid Fe at 1620° was found to be 0.0025 at 1 atm pressure. The same const. was obtained over a range of 0.01-0.04% C and 0.003-0.10% O. Expressed as the product of $\frac{1}{100}C$ and $\frac{1}{100}O$ it becomes 0.011. H. C. PARISH.

Improving castings with the double forehearth. CARL REIN *Gieserei-Ztg.* 27, 8-11(1930)—Cast iron melted in a cupola can be improved by a complete deslagging. This is accomplished by allowing the melt to stand a while so that the slag particles may rise to the top. Where various melts are melted in the same furnace the double forehearth has the advantage of keeping the individual melts separate. It is heated to a white heat with an oil burner before tapping, so as to not cool the charge. C. L. W.

Development of the modern blast furnace. GEO. F. ROSE *Blast Furnace & Steel Plant* 19, 255-7(1931) E. H.

The cupola furnace and its main dimensions. LEONOLD SCHMID *Gieserei-Ztg.* 26, 567-77(1929)—A discussion of the construction and operation of the Fe-foundry cupola, with and without forehearth. CURTIS L. WILSON.

The combustibility of coke in cupola practice. J. GWOSDZ *Gieserei-Ztg.* 26, 578-82(1929)—A dense, hard, difficultly combustible coke is the most suitable for cupolas, as it does not reduce CO_2 back to CO so readily. CURTIS L. WILSON.

The selection of coke for foundry purposes. OTTO HERRMANN *Gieserei-Ztg.* 27, 35-7(1930)—Low H_2O , ash, and S content and high strength are desirable. For cupola coke, low activity (velocity of reducing CO_2 to CO) is debatable. This activity is attributed to the unoxidized amorphous C atoms. When activity is desired, the graphitization caused by coking at too high a temp. is to be avoided. Cf. C. A. 24, 221. CURTIS L. WILSON.

Wrought iron, some pros and cons, and the need for research. H. W. GILBERT *Metals & Alloys* 2, 25-30(1931)—A correlated abstract. A. J. MONACK.

Producing hydraulic cylinder castings. E. LONGDEN *Iron and Steel Ind. and Brit. Foundryman* 4, 115-8(1931). C. H. LORIO.

Character of change of hardness of iron ruptured by tension. M. MIKHAILOV AND I. CIURBANOV *Acta Univ. Arad. Medice, Series XI, Tech.* No. 3, 1-5(1929)—Hardness tests (Shore scleroscope and Rockwell) made on samples of iron broken by tension showed that the hardness of the surface and of deeper layers increased in all parts of the specimen except the necked portion. On the portions corresponding to the neck the hardness of the surface layers decreased 21%, but the hardness of the inner layers increased 20%. B. N. DANILOFF.

The strength of gray cast iron at elevated temperatures. J. W. DONALDSON *J. West. Soc. Iron Steel Inst.* 38, Pt. II, 11-20(1930)—Cast irons comprising 2 plain irons contg. 0.65 and 1.48% Si, resp., and alloy irons contg. (a) 2.43% Mn, 0.40% Cr, 0.75% Ni, 0.12% V, 0.48% W and (b) 0.61% Cr, 0.30% Ni, 0.02% V, resp., were investigated to det. (1) the temp.-stress curves of a series of plain and alloy irons, (2) the influence of previous heat treatment on the temp.-stress curves, and (3) the effect of low Si and various alloy additions on the strength of cast Fe at elevated temps. both as cast and after heat treatment. DOWNS SCHAAF.

Plastic and elastic deformation. E. KIEFT *Iron Steel Eng.* 8, 54-60(1931).—The compression resistance of wrought iron and mild steel increases to unknown values when we compress plastic material. The rate is very great when the height is very small as compared with the diam., because of the rapid increase of expansion or horizontal resistance. An observed increase in resistance per sq. in. was from 50,000 to 230,000 lb. The applied force works vertically and is transmitted by deformation or displacement of the molecules into horizontal direction. Expansion work must be considered along

with compression work when deriving a formula in which tensile strength or compression strength of steel is taken as a constant during the whole operation. The solution of problems relating to plastic and elastic deformation involves application of the law of least resistance or least work. Twenty-two equations are evolved relative to the plastic flow of hot steel, of cold steel and of elastic flow. The strength can be improved, and a considerable amount of material can be saved by proper design of frames, rollstands or similar structures. Application of the theory of least work, credited to Castigliano, yields the same result as indicated in the formulas. W. H. ROYTON.

Wearing tests on iron and steel. ARVID JONASSEN. *Jernkontorets Ann* 114, 551-71 (1931) — A critical discussion is given of various tests for wearing resistance, including the Robin, Brinell, Spindel and Amder tests. The general conclusions are that wearing resistance increases with increased C content, with increased coarseness of structure and with increased speed of rotation. The general status of the results is, however, highly uncertain. Problems for further research are the effect of crystal structure as deduced by chemical analysis, heat treatment, final working temp., and slag inclusions. Seven references are included. H. C. DICK.

Investigation of a raw steel from pre-Roman times. H. HANEMANN. *Stahl u. Eisen* 51, 67-8 (1931) — A steel bar from the pre-Roman La Tène period was subjected to chemical and metallographical analyses. The irregular structure showed dark parts consisting of pearlite and cementite and lighter parts consisting of ferrite and pearlite. The chemical analysis showed 0.44% C, 0.065% Si, 0.10% Mn, 0.042% P, 0.012% S, traces of Cu and no other metals. The steel of that period could only have been made in the bloomery hearth. Forging experiments showed that sharp knives could be made from this steel but not without some difficulty due to the formation of cracks. This steel bar, over 2000 years old, is a good example of the great stability of cementite. J. A. SZILARD.

Testing method for the determination of the surface characteristics of [steel] plates. E. GEROLD. *Stahl u. Eisen* 51, 104-6 (1931) — The luster of the surface of iron plates is a measure of their smoothness. Several methods to measure the luster of plates are described. The polarization luster measuring apparatus built by Schmidt and Haensch (cf. K. Kiewer 2. *angew. Chem.* 32, 357 (1919)) utilizes a polarization photometer to measure the amount of polarized light in rays reflected from the surface to be tested. The stray luster measuring device, built according to Kempf-Flügge by the same firm, photographs the image of a small slit by means of the rays reflected from the surface to be tested. The clearer the image of the slit, the better the reflecting power and the smoother the surface of the plate. The Aclania luster-measuring app. is built on the same principle except that instead of exposing a sensitive paper, the observer reads the luster of the surface under test directly by comparing the divergency of the rays from the same point of the surface. The app. designed by the research lab. of the Verringte Stahlwerke Dortmund is similar to the above except that a photoelectric cell is used to measure the amt. of light reflected from the surface. Because of the comparatively small surface area possible to cover with the above app. this method is not suitable as an acceptance test for plates. However, it gives valuable information on the surface characteristics of plates when the adhesion of paints and varnishes is being studied. J. A. SZILARD.

Influence of high-frequency oscillations on the treatment of metallurgical products. G. MAHOUX. *Compt. rend.* 191, 132-30 (1930) — All expts. were made on rods 15 mm. in diam. Steel contg. 3% Ni, 3% Cr and 1% Mo was heated 9 hrs. at 500° in a current of NH_3 with a blank not electrically treated. The Brinellickers no. was raised from 340 to 1030, the resilience from 9 to 11.8 kg. and the nitride penetration was 0.35 mm. The only change in the blank was a nitride penetration of 0.01 mm. The B.V. no. of steel contg. 8% Ni, 19.8% Cr and 4% W treated 10 hrs. at 500° in NH_3 , was raised from 235 to 1035 at the surface, to 570.1 mm. beneath the surface and to 472 at the center. A soft steel electroplated with Cr and treated 9 hrs. at 530° showed a Cr penetration of 0.35 mm. the blank showed no penetration. Comments on the preceding communication. LEON GUILLET. *Ibid.* 1931-2. — The results have been confirmed by Aubert and Duval for steel and many alloys. A Ni coating on steel is absorbed at 450°. The treatment opens a fertile field for research. J. H. M.

The formation of cracks in steel castings. K. SINGER AND H. BENNEK. *Stahl u. Eisen* 51, 61 (1931) — The possibilities of the formation of so-called hot cracks during the cooling of steel castings and the filling up of these cracks or fissures by the liquid part of the casting are reviewed with reference to practical examples. The cracks are caused by great temp. differences in the casting by the manner of heat transfer from one part of the casting to the other and mainly by the size and shape of the coldest

part of the casting, from which during solidification contraction forces are acting on the rest of the casting. There is no direct relationship between piping and cracks, although by working toward a pipe free casting the better temperature conditions prevent the formation of cracks and help in filling up those formed by the liquid part of the casting during the cooling period. The importance of the proper contraction of the casting is emphasized. Were it possible to control the shrinking of the casting the danger of the formation of cracks would be greatly reduced. It is pointed out that in a casting solidifying in a field of a high frequency coil the induction current would counterbalance the radiation losses thereby preventing the formation of cracks. J. A. SEILARD

Aging of steel castings. A. FOUR. *Stahl u. Eisen* 50, 449-1 (1930).—The impact strength of mild steel castings (notched bars) containing C 0.1-0.27, Si 0.16-0.16, Mn 0.45-1.37, P 0.005-0.006 and S 0.01% decreases on aging at 250° for 1 hr. after forging to a reduction in area of 7-10% but not nearly to the same extent as it does after completely annealing. The values obtained for the annealed metal at 20° to -20° were consistently below 2 m. kg./sq. cm. whereas those for the aged metal varied from about 6 m. kg./sq. cm. at 20° to about 2.3 m. kg./sq. cm. at -20°. B. C. A.

Effect of various annealing temperatures on cold-worked low-carbon steel. H. E. PETERLOW. *Metals & Alloys* 2, 14-11 (1931). The results are represented by 3 graphs which show (1) the relation between annealing temps. and Rockwell hardness of 4 low C steel brake drums, (2) the relation between annealing temps. and ductility and (3) the relation between annealing temps. and hardness and ductility of cold-worked low C steel. A. J. MONACK

Cold-rolled steel. GEORGES DELBART. *Aciers speciaux* 5, 483-78 (1930).—This work covered (1) influence of cold rolling and successive annealings on the tensile strength and elongation, (2) hardness, (3) corrosion by H_2SO_4 , (4) attack by fuming H_2SO_4 , (5) influence of cold rolling and successive annealings on the impact strength and (6) influence of cold rolling and rapid annealings on the magnetic permeability. Data and graphs are given in connection with each of the items above. A large number of photomicrographs is included. A. J. MONACK

Steels for cold pressing. G. R. BOLSOVER. *Iron and Steel Inst. and Brit. Foundryman* 4, 101-2 (1930).—Low C running steels used for deep drawing do not work harden readily. The tensile strength varies approx. as one third the % reduction in steel thickness. When steel is stretched more than 5% brittleness is encountered on reheating to 200° to 300°. Age hardening after quenching increases the tensile strength of dead mild steel 67%. For deep drawing purposes stainless steel must be low in C and should be quenched from a high temp. C. H. LORIE

Tensile strength at high temperatures of steel containing small quantities of nickel and molybdenum. W. LIESMANN AND C. SALTMANN. *Stahl u. Eisen* 50, 442-6 (1930).—The tensile properties of steel with C 0.2, Mn 0.8, Si 0.2%, Ni 0.4-2.2 and Mo 0.2-0.6% have been determined at 500° after annealing at 900-950°. Ni and Mo both increase the tensile strength and yield point but reduce the ductility, 0.1% Mo increasing the ultimate strength by 1 kg./sq. mm. and the yield point by 2 kg./sq. mm., and 0.5% Ni increasing the former by 2 kg./sq. mm. and the latter by 1.8 kg./sq. mm. With higher proportions of Mo the ratio ultimate strength/yield point is unfavorably affected, but further addition of Ni improves this ratio. The steel with 2.1% Ni and 0.5% Mo has a yield point of 27.5 kg./sq. mm., a tensile strength of 43 kg./sq. mm., an elongation of 11% and a reduction in area of 20% at 500°. B. C. A.

Solution of cementite in carbon steel and the influence of heterogeneity. ERNEST WALLELOW. *Engineering* 131, 27-30 (1931).—Transformation of α iron to γ iron starts from the grain boundaries. Com. steels show a transformation interval, the transformation starting at different temps. in different parts of the specimen. A new material, "sorbo troostite," is found when heating at 3.5-5.2° above A_{c1} for 1 hr. or less. Its structure is less distinct than that of sorbite and it is not so strongly attacked by HNO_3 as troostite. It appears to be residual undissolved ferrite and cementite and never to have been in the γ state. It is found in the peripheral parts of austenite grains. Varying soly of pearlite grains is not due to alloying elements as contaminants but to their irregular occurrence in the structure. ANN NICHOLSON HIRD

The development of nitriding plants. F. BRUHL. *Krupp Monatshefte* 11, 287-90 (1930).—A description is given of the development of nitrided steel together with a discussion of the requirements demanded of nitriding equipment. A number of installations are described. H. STOEHRZ

The case-hardening of commercial steels by means of nitrogen. B. JONES. *Iron and Steel Inst. and Brit. Foundryman* 4, 77-86, 99 (1930).—Five British com. austenitic steels of various compns. were nitrided at 500° for 90 hrs. Except for a 13.20% Mn

steel with 1.25% C, the increase in hardness was inappreciable. The austenite in the Mn steel was partially decomposed into troostite-martensite. Nickeliferous austenite does not harden on treatment. The effects of certain elements were ascertained by nitriding 20 com. steels of different chem. compns. for 90 hrs. at 500°. Contrary to published information, plain C steels are nitrided. Mild steels harden to a greater extent than do high C steels. Nickel steels do not nitride. With Cr and V the retarding effect of N₂ is not so great. The effects of the combination Al, Cr, Mo appear to be more than additive. C. H. LORIG

Use of zirconium alloys in steel manufacture. *Proc. Roy. Soc. London* 135, 509-21 (1930).—The following are discussed: (1) action of Zr on the impurities contained in steel; (2) phys. properties of Zr steels; and (3) steels for which the use of Zr is recommended. Data and photomicrographs are given. A. J. MONACK

New manganese-silicon alloys for the deoxidation of steel. C. H. HERTY, JR., AND G. R. FITZGERALD. *Bur. Mines, Rept. of Investigations* 3081, 14 pp. (1931).—By using Mn-Si alloys for deoxidation, larger inclusions were formed than by using Si alone. The av. size of the inclusions formed when using Si was 15×10^{-4} cm. The size increased with increasing Mn, until at a Mn/Si ratio of 4/15 the largest particles in dead-killed steels were $1000 \text{ cm} \times 10^{-4}$. These large particles would rise at a much more rapid rate (600-2400 in. per min.) than the small SiO₂ particles (0.15-1.5 in. per min.). Tests in commercial acid and basic open hearths show the same results on steels of from 0.12 to 0.60% C using alloys with a Mn/Si ratio of 4.5/0.0. Mn efficiencies of 75-83% have been obtained when forging grade steels have been deoxidized with these alloys. ALDEN H. EMERY

Some important advances in the properties and treatment of metallic alloys. LEON GUILLET. *Rev. metal.* 27, 449-66 (1930), *Rev. univ. mines* 4, 221-30, 296-303. — General review. H. C. DAVIS

Examination of the resistance of metal alloys against attack by salt solutions. III. L. MAASE AND W. WIEDERHOLT. *Korrosion u. Metallschutz* 6, 265-77 (1930), cf. C. A. 25, 1205.—Sols. of sylvite, carnallite, Glauber's salt, and bitter salt were used as corrosive media, in which the specimens previously specified were immersed under various conditions. The testing period was 4 months in the expts. carried out at room temp., and 8 days with those carried out at 90°. The losses of wt. and the appearance of the corroded surfaces and of the corrosion products are given in detail. The greatest losses of wt. were observed in the group of the various forms of iron. By addn. of Cr and Ni the material is considerably improved, and Cr-Ni steels give the best results that were observed. The second best group is the Cr-Ni-Fe alloys. Ni and the other Ni alloys show a similar behavior, although they are not quite as good. Between them and plain iron rank Cu and the Cu alloys. No essential difference is observable between Cu bronze and brass. The losses of wt. of the β brasses are slightly lower. The magnitude of the attack, generally speaking, is not so great in sols. as when the sheets are covered with solid salt. The corrosion is mainly restricted to the surface, without penetrating deeper into the metal. The penetration amounts to approx. 0.02 mm. in 4 months. Perforation of the sheets by pits was not observed. A soln. of sylvite was found to possess the greatest corrosiveness. Generally speaking, the effect of Glauber's salt was much smaller and that of carnallite even less, although exceptions were observed in the series where O₂ was passed through the sols. The influence of the experimental conditions was considerable. The losses of wt. were doubled by using the double quantity of corroding soln. per sq. cm. of area. The same holds true if there is a boundary line of liquid and air. Introduction of O₂ causes losses 5 times as high. The corrosion at 90° was found to be greater after 8 days than that at room temp. after 4 months. LEOPOLD FESSLER

Structure analogies of intermetallic phases. A. WESTGREN AND W. EKMAN. *Arkiv. Kemi, Mineral. Geol.* 10B, No. 11, 1-6 (1930).—In intermediate phases of the β brass type contg. either Cu, Ag, or Au, it has been found that there is a ratio of valency electrons to atoms of 3/2, while in the lattices of the γ brass type the ratio is 21/13. Recently metallic compds. have been found not contg. Cu, Ag, or Au with atomic groupings resembling β or γ brass, viz. NiAl, NiAl₂, and FeAl, where the atoms of the transition elements behave as if they had a valency of zero. To test the validity of this remarkable rule, the authors made a series of alloys corresponding to the general formula A_1B_n , where A is a transition element (Co, Ni, Rh, Pd, Pt), and B is either Zn or Cd. The alloys were made in sealed evacuated quartz tubes by heating for about a week at 900-1000°. In all cases where alloying succeeded, a homogeneous material of γ -brass structure resulted, as was evidenced by the powder photograms. The homogeneity range extends for Co-Zn alloys from 15 to 22 at. % Co and for Ni-Zn alloys from

15 to 19 at. % Ni. It was noticed that alloying of transition elements with Al, Zn or Cd gives a much more pronounced decrease in vol. than is the case with alloys formed with atoms having complete inner electron shells. The contraction for Ni-Cd amounted to almost 8.7% although Ni occupies only 19 at. % in the alloy. H. S. V. K.

Packing of atoms in alloys. A. WERTSCH and A. ALMAY, *Z. physik. Chem.*, Abt. B, 5, 14 (1929).—The variation of the linear lattice dimensions with composition of the alloy has been investigated in the α phases of the systems Ag-Al, Cu-Al and Ni-Al and in all cases the parameters are less than those calculated by Vegard's additivity law from the parameters of the components. From calcs. of the vol. occupied per atom in the various phases of the above systems, and of the systems Cu-Zn, Ag-Zn, Ag-Cd, Cu-Mg and Fe-W, it is shown that in general, a contraction in at. vol. accompanies the formation of the intermediate phases. The at. vols. in contiguous phases of any system differ only slightly so that the curves of vols. drawn for a series of phase regions are almost continuous. The vol. contraction is especially marked in Ni-Al alloys, and it is suggested that this depends on the fact that such alloys are formed by the union of an atom which is deficient in a nuclear electron with one of normal type. B. C. A.

The influence of temperature upon the notch toughness of aluminum alloys. TORAO KOBAYASHI, *Mem. Res. Lab. Eng.* 3, 145 (1930).—Duralumin, V-alloy and other com. Al alloys were worked, quenched and aged 20 days, after which test pieces were placed in an impact machine and tested at temp. intervals of 50° up to 550° . The heating arrangement was attached to the testing machine and could be removed simultaneously with the release of the pendulum. The notch toughness was found to resemble plasticity, in that it increased with increasing temp., reaching a max. at 470° . C. E. L. WILSON.

Volume change during solidification of manganese and some alloys. KOTARO HONDA, YOSHIMARU MATSUYAMA and YATSUGI ISONO, *Bull. Inst. Phys. Chem. Research* (Tokyo) 9, 905-10 (1930). (*Abstracts* 88 (in English) published with *Sci. Papers Inst. Phys. Chem. Research* (Tokyo) 14, Nos. 275-6). Vol. changes during solidification of Mn and some alloys containing Mn, Ni or Co are measured by using the thermobalance. Values of V/V_0 (V = vol. of liquid sample at m. p., V_0 = contraction by solidification) obtained are: pure Mn (Mn 99.9%, C 0.01%, Al 0.03%, Fe 0.03%, Si trace), -1.50; Mn-C (C 1.10%), -1.60; Ni-C (C 2.23%), +2.05; Co-C (C 2.20%), +2.03. K. KONDA.

Macro-etching of tin-base bearing metals. ERNEST B. DRAKE, *Metals & Alloys* 2, 20-4 (1931).—A method for macro-etching tin-base babbitts of various composition is suggested. The failures in connecting rod bearings which have been run in test engines appear most often in the finely crystalline babbitts near the joint. The macro-etched babbitts indicate fairly well the size and distribution of the micro-constituents which may be expected. Centrifugally cast main- and connecting rod bearings show SnSb cubes, which are generally associated with a smaller no. of CuSn needles than when the SnSb cubes are absent, all imbedded in a groundmass of Sn-rich eutectic when the babbitt is of the type: 85% Sn, 7.5% Sb, 7.5% Cu. The die-cast main- and connecting rod bearings show no SnSb cubes but do show a larger amt. of CuSn needles imbedded in a groundmass of Sn-rich eutectic and delta solid soln. when the babbitt is of the type: 85% Sn, 7.5% Sb, 7.5% Cu. A. J. MONACK.

Production of lead-tin alloys from tin slag. EDMUND R. THREWS, *Metallics* 21, 3-4 (1931).—Sn slags are melted with Pb or Sn-poor alloys. A purer product is obtained by using type metal. Pb residues can be used with Sn slags to produce an alloy of 8% Sn. For alloys richer in Sn this product is oxidized and used as raw material for the next smelt. Sulfurous ores require a preliminary roasting. The presence of much As is harmful because of the formation of viscous iron arsenides. A coking mixt. containing 10-25% wood charcoal is best. ANN NICHOLSON HIRSH.

Zinc and zinc-alloy use in the automotive industry. II. ROBERT M. CURTIS, *Am. Metal Market* 38, No. 22, 5, 10 (1931).—This installment deals with Zn wire, extruded Zn shapes, die-castings, the use of fully automatic machines, the introduction of high grade Zn in die-castings, Zn alloys for body hardware, and other equipment. Some data are given on the Zn-Al alloy system. III. *Id.* No. 23, 5.—This installment deals with heat treatment of Zn-base die-casting alloys, water-vapor tests, the disastrous effects of contamination, phys. tests on die-casting alloys, and the use of inserts in these alloys. IV. *Id.* No. 24, 5, 10.—Cleaning, etching and Ni plating are discussed. A series of tests indicates that Cd plate may be applied directly to die-casting with no appreciable harmful effects. Attractive coats of enamels, japans and lacquers may be applied to Zn and its alloys. Metallic Zn is used in the automotive industry.

for electro-galvanizing as a constituent of brass, as rolled sheet and strip for a variety of fabricated products and in the production of a great no of die-cast parts. The phys. properties of the various Zn alloys and the ease of application of attractive and durable plated finishes have resulted in annual increases in Zn consumption for automobile parts. Tables show the A. S. T. M. specification B 6-18 for Spelter and the phys. and elec. properties of Zn die-casting alloy contg. 4% Al, 3% Cu, 0.03-1.0% Mg and the balance Zn. W. H. ROYTON.

Fatigue strength of unwelded and welded cast and rolled materials. WALTER B. BARRELS. *Z. Ver. deut. Ing.* 74, 1423-6, *Gusserei-Zig* 27, 607-10, 637-43, 661-4 (1930).—The type of graphite sepn. is of special significance for fatigue strength. In fine eutectic form a considerable increase in vibratory strength is obtained, even with welded bars. The skin produced on the surface of castings by melting of quartzite is of little significance for increased static or dynamic strength, but the grain structure of the outer zone is important. Annealing in all cases lowers the fatigue resistance of cast iron. The effect of welds upon low C steel was studied. The dynamic strength of the weld was greatly influenced by the dynamic strength of the materials welded. The weld in general has a stiffening effect, and hammered welds increase resistance to fatigue. H. STOERTZ.

Building up metals with electric arc welding. K. TEWES. *Gusserei-Zig* 27, 553-9 (1930).—In arc welding it is customary to add materials to the electrodes which will improve the quality of the weld. Si and Mn have been widely used to bring about deoxidation of the weld, but Ti appears to be more suitable. It combines energetically with O and materially reduces the quantity of oxides in the weld. The TiO_2 which is formed readily floats to the top and can be removed with the slag. Greater homogeneity of the weld results in improved mech. properties. Ti also tends to form nitrides, Ti_3N , Ti_2N , TiN , and Ti_4N_3 , which greatly increase the hardness of the steel. H. STOERTZ.

Soft steel welds deposited by the electric arc. D. ROSENTHAL AND M. MATHIEU. *Compt. rend.* 191, 484-6 (1930).—The nature of the weld of soft steel bars produced by the elec. arc depends on the manner in which the molten metal is protected from oxidation. Metallographic examn. shows that welds protected with a covered electrode have more regularly formed grains than non-protected welds. X-ray diagrams show the existence of strains in non-protected welds. I. G. VANDER BOSCH.

Non-corroding and acid-resisting alloys. BO KALLING. *Tek. L'ér'lad* 75, 101-3 (1928).—A review. C. A. ROSE.

Some properties of protective films on metals. ERNEST S. HEPBURN. *Chemistry and Industry* 50, 21-5 (1931).—A review of film properties is given, with special consideration of the work carried out by Evans and by the author. A theory of anodic polarization is offered, according to which OH^- ions are discharged simultaneously with other anions, such as SO_4^{2-} ions for instance, if a certain crit. c.d. is reached and more SO_4^{2-} ions are discharged than can be replaced. This has the effect of producing a layer of oxide or basic salt on the surface. The presence of SO_4^{2-} ions renders this film non-protective, but it has the effect of raising very greatly the effective c.d. at those areas of the electrode which are uncovered. The consequence will be that in some spots the discharge will consist of OH^- ions only, which give a perfectly continuous oxide film. The dis. continuous film first formed will be thrown off the electrode by the subsequent evolution of O_2 . LEOPOLD PFESSEL.

The mechanism of the suppression of corrosion velocity by colloids. W. BECK AND F. V. HESSENT. *Z. Elektrochem.* 37, 11-20 (1931).—The decrease in corrosion velocity of Fe in H_2O or strong acid solns. is due to the presence of hydrophobic colloids, and is dependent upon the type as well as the concn. of added colloid. The authors suggest the possibility of the formation of coagulation layers upon the surface of the metal due to the action of the Fe ions upon the negatively charged particles of the hydrophobic colloid. The colloidal materials studied were agar, starch, silicic acid and gelatin and the metals. Arminco' electrolytic, carbonyl and pure Fe. H. STOERTZ.

Protection of large steel structures from rust. HENRY E. WEITKAMP. *Het Gas* 50, 450-4 (1930).—A review. B. J. C. VAN DER HOEVEN.

Initial corrosion rate of steels. H. O. FORREST, B. E. ROSTHELI AND R. H. BROWN. *Ind. Eng. Chem.* 22, 1197-9 (1930).—Studies of corrosion rates in distd. water, of oxide-free, low C 14% Cr steel and of (18-8) Cr-Ni steels were made during the initial stages of corrosion. After 10 min. the rates for the Cr and Cr-Ni steels became essentially zero, while the rate for C steel decreased but slightly, indicating the formation of invisible, adherent, protective films on the resistant steels. A method for obtaining reproducible results is described. B. E. ROSTHELI.

Rust-free steels, their properties and preparation. E. HOUDEMENT. *Krupp Monatsch* 11, 270-86 (1930).—H discusses the various types of stainless steels as classified under the headings martensitic, ferritic and austenitic steels. The system Fe-Cr-Ni is studied together with the effect of C upon it. A table is given giving the range of C, Si, Mn, Cr, Ni, W, Mo and Cu in typical examples of these steels. The prep'n of 1.2 A in a Siemens Martin furnace and the prep'n of a Cr steel by reduction of a Cr ore in a basic furnace are described. Heat treatment, hot and cold working and recryst'n are considered in relation to their effects upon mech. properties. H. S.

Behavior of rustless steel toward dilute sulfuric acid. J. FARR. *Krupp Monatsch* 11, 267-9 (1930).—Conclusions drawn by various investigators that rustless steel is attacked by H_2SO_4 are incorrect. If practical requirements are adjusted so that the passivity of the rustless steel can be maintained, it is possible to use it in many solutions free H_2SO_4 . H. STORITZ.

Influence of a small addition of copper on the corrosion resistance of structural steel. O. BALER, O. VOGEL AND C. HOLTHAUS. *Mitteil. der deut. Materialprüfungsanstalt Sonderheft* XI 25 pp. (1930). Extended abstract by V. A. KENDALL in *Metals & Alloys* 1, 890-6 (1930). The samples covered a range of soft steels, structural steels and 0.25 to 0.35% C steels. The tests were conducted as follows: (A) at room temp., (1) in dist'd H_2O , (2) in Spruce River H_2O , (3) in North Sea H_2O , (4) in humic acid sol'n, (5) alternating wet and dry in (a) 1% NaCl sol'n and (b) dist'd H_2O , (6) in the atm., (7) in 1% H_2O_2 and (8) in 1% HCl. (B) at temp. of 50°, (1) in dist'd H_2O , (2) in Spruce River H_2O , (3) in North Sea H_2O , (4) in humic acid sol'n and (5) in 1% H_2SO_4 . Data and graphs are given. A. J. MONACK.

Corrosion of metals by phosphoric acid. PETER R. KOSTING AND CONRAD HILLY, JR. *Ind. Eng. Chem.* 23, 140-50 (1931).—Tables are given of corrosion rates of 22 metals and alloys with pure and crude phosphoric acid of various strengths. Notes on passivity and corrosion inhibitors present in H_3PO_4 are added. ANN. NICHOLSON HILL.

Furnace for metallographic examination of specimens at high temperatures. B. A. ROGERS. *Metals & Alloys* 2, 9-12 (1931).—The furnace is described. Expts. with the furnace have shown that it is possible to observe marked surface changes upon some elements as these pass through certain allotropic transformations. Observations at a magnification of 50 to 75 diams. were conducted on samples heated above 1500° and at higher magnifications when the temp. did not exceed 1000°. Photomicrographs were taken at magnifications of about 150 diams. for temps. up to 1000°. Ten photomicrographs are given. A. J. MONACK.

Investigations on wearing in the Amsler machine. C. J. GUMMAR MALMBERG. *Jernkontors Ann.* 114, 572-62 (1930).—Ring material for trolleys and railway coaches was used. Aside from the quality of the material, the amount of wear is shown to be dependent on (1) the total load, (2) the local strains, (3) the slippage and (4) the wearing distance. Standardization requires fixation of load (P), diam. (D), r. p. m. (n), and the time (T). In the actual runs, the test cylinders (60 mm. diam.) were turned out of wrought iron, while the whetting cylinders (60-65 mm. diam.) were punched from rolled rods which were not heat treated after the rolling. The diam. of the rolls was essentially of no influence between 35 and 50 mm., with 10% positive slippage, but the wear was only about 25% as great for 50 mm. as for 35 mm., with 10% negative slippage. Increasing C from 0.6 to 0.8% quadrupled the wear, but the min. wear was with 0.90%. The wear also increases with the slippage. Wear on the test cylinder and the whetting cylinder tends to run parallel. The wear decreases with increased Brinell hardness, and the total wear is a min. when the hardness of the test cylinder exceeds that of the other cylinder. Oil hardened test cylinders were down the whetting cylinder much more than did water hardened test cylinders. Structure plays a large part in wearing resistance. No great difficulty should be encountered in standardizing dimensions of the test piece, load, slippage, r. p. m. and duration of test. H. C. D.

The relation between the H overvoltage and composition of brass (THOMPSON) 4. The electrolytic Zn plant of the Evans-Walloway Company at East St. Louis (TAINOR, BOSQUT) 4. 1930 steel plant electrical developments (KENYON) 4. An air-thermostat for corrosion research (EVANS) 1. Annual review number [of *Eng. Mining J.*] (GRANT, et al.) 18. Decorating [metal surfaces] (Brit. pat. 335,788) 19. Recovering dry palm oil from its admixtures with water [formed by washing palm oil from tin plate] (U. S. pat. 1,790,748) 27. Rotary kiln for burning ores (U. S. pat. 1,791,282) 20. Fuel briquets [from ores] (U. S. pat. 1,791,077) 21.

Aluminum in Aircraft. Pittsburgh: Aluminum Co. of America. 150 pp. 50 cents. Reviewed in *Metals & Alloys* 2, Book Review Sect., 53(1931).

GROSSMAN, M. A., AND BARN, E. C. High Speed Steel. New York: John Wiley & Sons. 173 pp. Cloth, \$3.50. Reviewed in *Metals and Alloys* 2, Book Review Sect., 55, *Ind. Eng. Chem.* 23, 341(1931).

Ingot Contour and Its Relation to Sound Steel. Baltimore: Gathmann Engineering Co. 90 pp. \$1. Reviewed in *Metals & Alloys* 2, Book Review Sect., 53(1931).

Iron and Steel Industry Handbook and Directory. London: Louis Casser Co., Ltd. 234 pp. 10s. Reviewed in *Metals & Alloys* 2, Book Review Sect., 54(1931).

Jahresbericht der Abteilung für Metallchemie und Metalltechnik. Reprint from "Jahresbericht VIII." Berlin: Verlag Chemie. 54 pp. Paper, M 6.40. Reviewed in *Metals & Alloys* 2, Book Review Sect., 51(1931).

Metal Industry Handbook and Directory. London: Louis Casser Co., Ltd. 400 pp. Not sold, given to subscribers to *Metal Industry*, London. Reviewed in *Metals & Alloys* 2, Book Review Sect., 51(1931).

Metallurgie Agenda Dumod, 1931. Paris: Dumod. 497 pp. Cloth, F 23.50. Reviewed in *Metals & Alloys* 2, Book Review Sect., 54(1931).

SCHULT, W. Ferrosilicium. The Hague: Algemeene Landsdrukkerij. 47 pp. STERNER-RAISER LEONIE; Die Edelmetall-Legierungen in Industrie und Gewerbe Leipzig: Dehner. 154 pp. M. 8.75.

Vanadium ores. ALPHONSE GRIMMINGER and ÉMILE COMPAGNE. Fr 632 654, June 24, 1929. Ores of V, particularly vanadates of Pb or other metals, are submitted in the presence of a suitable flux, to a regulated reducing action to obtain the heavy metals in the metallic state and the V in the form of the iron trisulfide.

Treating rare metal silicate ores. CHARLES V. LEBLANC (to Westinghouse Lamp Co.) U. S. 1,791,272, Feb. 3. An ore such as rhenium is heated with C to a sufficiently high temp. (suitably about 1700-2500°) to convert the Si and rare metal content into carbidic compds. which are sepd. by dissolving the Zr carbidic with dil. HNO₃.

Recovery of precious metals from solutions. H. WOLFRUM. Brit. 235,555, Feb. 2, 1929. Acid solns. contg. precious metals such as chlorides of Pt, Pd, Au or Ag obtained from ores and concg. also base metals such as Cu, Fe or Ni are brought into contact with active C, which adsorbs the precious metal compds. Strongly acid solns. are first dil. or partly neutralized. The adsorbed compds. may be recovered by the action of concd. HCl.

Vacuum and pressure filtration apparatus suitable for metallurgical pulps, etc. JOHN T. SIMONIN. U. S. 1,791,251, Feb. 3. Structural features.

Preparation of oxide ores for reduction. V. J. J. DEFAVE. Belg. 370 110, June 30, 1930. A gas rich in CO is passed over the ore, heated to a suitable temp. to produce a deposit of C, which facilitates the subsequent reduction.

Zinc blende treatment. SOC. ANON. LA NOUVELLE MONTAGNE. Brit. 325 600, March 9, 1929. Zn blende, either raw or roasted, is entered under forced draught in admixt. with ZnSO₄ in the proportion of about 3 mols. sulfate to 2 mols. blende in excess of the amt. of sulfide necessary to ensure combustion and a fluid such as gum or dextrin may also be added. Cf. C. A. 24, 3012.

Metallurgical furnace suitable for heat-treating metal sheets, bars, packs, etc. FRANK A. FAHRENWALD. U. S. 1,791,494, Feb. 3. Horizontal shafts pass through the side walls of the furnace and actuate circular disks and eccentric conveyor rolls.

Supporting frame for metallurgical furnaces. DIJONLENGRE MAISONNETFABRIK A-G. Fr 636,045, Mar. 29, 1930.

Rotary furnace for ore treatments such as roasting or reducing with gases. CARL P. DIERCK. U. S. 1,791,493, Feb. 3.

Improvement in the construction and operation of blast furnaces. V. J. J. DEFAVE. Belg. 370,554, June 3. The air is blown into the furnace at a pressure considerably above atm. The direction of the air blasts is varied, either continuously or intermittently by appropriate mounting of the nozzles.

Blast furnace, etc., operation. VULCAN FURNACES A-G. Ger. 515,215, Sept. 6, 1927. The introduction of water into the combustion zone of blast furnaces, cupolas and other shaft furnaces is effected with the use of a solid, liquid or gaseous auxiliary substance, other than the blast, as a carrier. Coke is specified.

Gas burner for blast-furnace stoves or furnaces. WM. R. PAULSEN (to American Heat Economy Bureau). U. S. 1,791,911, Feb. 3.

Drying air for furnace blasts. GÖTTFRID MALYOTH. Ger. 515,216, Oct. 5, 1929.

The air is led over CaC_2 . The mixt. of air and C_2H_2 , thus obtained may be preheated and supplied as such to the furnace, or the C_2H_2 may be burnt, thereby regenerating one half of the moisture initially present in the air. Cf. C. A. 24, 5013.

Tuyère for cupola furnaces. FRIEDRICH SCHINKE. Ger. 515,233, Mar. 21, 1929.

Bessemer process of making steel. FRANK W. DAVIS (to S. G. Allen Trustee to F. W. Davis). Can. 308,761, Feb. 17, 1931. The Bessemer process of making steel with non-Bessemer pig consists in detg. the net blow heat required and then proportioning the amt. of O to the amt. of inert gases in the blast so as to reduce the heat in the waste to a point which will give the blow heat required.

Casting ingots. T. BRINKMANN. Brit. 335,538, March 8, 1929. Mech. features.

Two-part ingot mold. SIEGFRIED JUNGHAUS. Ger. 514,360, Jan. 8, 1929.

Ingot mold and hot-top construction. H. J. DARLINGTON (to W. M. Charman). Brit. 335,480, Dec. 31, 1928.

Apparatus for melting metal waste. HIASCH, KUPFER- UND MESSINGWERKE A.-G. (Otto Götz, inventor). Ger. 514,442, Jan. 30, 1929.

Rare metals. FANSTEEL PRODUCTS CO., INC. Fr. 693,417, April 5, 1930. Difficultly fusible rare metals, particularly Ta, are freed from C which they contain as impurities by mixing the metal with a convenient amt. of an oxide of an element which volatilizes at a temp. lower than the m. p. of the metal, the mixt. being afterward heated under vacuum to a temp. at which the C oxidizes and at which the residue left by the oxide volatilizes. The oxide used may be MgO .

Sheet aluminum. VEREINIGTE SILBERHAMMERWERKE HETZEL & CO. Fr. 693,869, Apr. 14, 1930. Sheet Al which is resistant to corrosion is made by welding into one whole, 2 or more laminated plates of Al and afterward transforming into sheets by rolling.

Tin. SOC. D'ÉLECTROCHIMIE, D'ÉLECTROMÉTALLURGIE ET DES ACIÉRIES ÉLECTRIQUES D'UGIVE. Fr. 692,640, June 22, 1929. In extg. Sn from its ores, alloys or slags, an auxiliary element contg. a proportion of Si which allows the formation of an Fe silicide unattacked by HCl or by an acid soln. of SnCl_2 , is added to the initial charge.

Removing lead from metals. DUDZEELE CORP. OF AMERICA. Fr. 692,721, Mar. 25, 1930. Layers of Pb are removed from metal objects by immersing the objects in a warm acid soln. of one or more chlorides, e. g., a soln. of NaCl contg. H_2SO_4 or HCl or a soln. contg. FeCl_2 and NaNO_2 .

Lead containing phosphorus in small proportion. AMERICAN MACHINE & FOUNDRY CO. Brit. 335,545, March 19, 1929. A soldering, metal-coating or deoxidized lead is prep'd. with a content of P (which may be 0.01–0.03%) predet'd. in accord with the rate of oxidation and the O content of the metals with which it is to be used in contact.

Magnetic iron and steel. VEREINIGTE STAHLWERKE A.-G. Fr. 693,602, April 9, 1930. Fe or steel free from magnetic aging and to be used as magnetic material is made by lowering the absorption of O and (or) deoxidizing the steel in known manner.

Cementation of steel. H. J. SCHIFFLER. Ger. 514,479, Mar. 19, 1927. Absorption of C in an amt. above the eutectic ratio is hindered by using steel contg. 0.4–4% of Al.

Rustless steel. GEORG MÜLLER. Fr. 692,887, Mar. 23, 1930. A rustless steel contains Cr 0.1–1, Cu 0.25–0.5, Ni 0.04–0.25 and W 0.01–0.03% and is free from Si. The added metals are introduced in the form of an alloy with Fe contg. Cr 20–30, W 5–6, Ni 6–10, Cu 10–15 and C (present in the form of carbide) 2.5–3%.

Treating surfaces of metals such as iron in order to prevent rusting. EARL K. WALLACE (to Rust Prevention Chemical Co.). U. S. 1,790,960, Feb. 3. The surface is treated with a mixt. formed of pure H_3PO_4 , 42.5, water 53 and casein 1.5%.

System of magnetic testing to determine properties of railway car wheels, etc. CARL KINSLEY (to Magnetic Analysis Corp.). U. S. 1,790,819, Feb. 3.

Alloy. VEREINIGTE STAHLWERKE A.-G. Fr. 692,743, Mar. 25, 1930. Hard alloys of the stellite kind, i. e., having a basis of Co, Cr, W and C with or without Mo, Ni, Fe and Mn, are made by using over 2% of C in the form of graphite.

Alloys. CAMILLE CONTAL. Fr. 692,634, June 21, 1929. See Ger. 491,815 (C. A. 24, 2420).

Refractory alloys. SOCIÉTÉ ANON. DES ANCIENS ÉTABLISSEMENTS SKODA A PILZEN. Fr. 693,775, Apr. 11, 1930. Refractory Cr alloys are made by compensating the high C content by the simultaneous addn. of Al and Ti so that the critical points of the alloy are displaced above the max. temp. of use. Thus, an alloy may contain a max. of 40% of Cr, C 3, Al up to 15 and Ti up to 15%.

Refining alloys. I. G. FARBENIND. A.-G. Fr. 693,544, April 8, 1930. Alloys of Mg and Ce are refined by adding dehydrated CeCl_3 to the mixt. of salts used in the

refining, whereby any moisture in the salts reacts with the CeCl_3 and does not affect the Ce in the alloy.

Aluminum and its alloys. WILHELM NEUMANN. Fr 692,538, Mar 21, 1930. Al or its alloys is obtained by the electrothermal reduction of Al_2O_3 , clay, kaolin, bauxite, etc., the vapor tension of the Al being lowered by means of substances of high b.p. during the reduction operation. The Al and the added substances are sepd. by a phys. process and a distn. Metals of higher b.p. than Al or Si may be used.

Iron alloy containing copper, silicon and aluminum. BYRAMJI D. SAKLATWALLA. Fr 693,418, April 5, 1930. See U.S. 1,781,527 (C.A. 25, 66).

Iron-chromium-nickel-tungsten alloys. R. ROSENHAHN and C. H. M. JENKINS. Brit. 335,558, April 23, 1929. Alloys are described which contain over 40% Cr and Ni together (the ratio of the Cr to Ni being between 2/3 and 1/1), together with 1% of C or Si or both. The total Cr and Ni content may be about 50-55% and the W 4-20% or more.

Lead alloy. ROBERT J. SPENKEMER (to S. & T. Metal Co.). U.S. 1,791,148, Feb. 3. A tough, slightly hardened, corrosion-resistant alloy suitable for covering cables, pipes, etc., comprises Pb together with Ni 0.01-0.03, Ca 0.1-0.3, Sn 0.25-0.5 and Al 0.02-0.1%.

Magnesium alloys. I. G. FARRENING A-G. Fr 693,204, April 2, 1930. The properties of alloys having a high content of Mg are improved by a heat treatment which increases the sol. of the materials in the alloy, followed by a cooling strongly retarded by artificial means, particularly in the zone from 270-190°, the duration of the cooling in this zone being at least 3 hrs. Cl. C.A. 23, 678, 1214.

Optical mirrors from metals or their alloys. I. G. FARRENING A-G. Fr 693,816, April 12, 1930. Metallic mirrors of a high and uniform reflecting power for all wave lengths are formed by one or more layers of different metals or their alloys, which are chosen so that the low reflecting power of one constituent for a certain part of the spectrum is compensated by the high reflecting power of another constituent. Ag along with Al, Si or Ni is mentioned as suitable.

Salt baths for annealing metals. W. FREY & Co. Ger. 514,276, Jan. 31, 1930. The baths consist of, or contain, NaCl, K_2SO_4 , and Na_2CO_3 (or K_2CO_3) in such amounts that the ratio Na:K is less than 1. The baths may contain also borax or Na_2HPO_4 up to 6, alk. earth oxides or carbonates up to 10 and NaNO_3 0.1-0.2%. An example given is NaCl 28, K_2SO_4 41.5, K_2CO_3 21, borax 4.5, MgCO_3 2.5 and CaCO_3 2.5%. The baths are particularly intended for objects of noble metals or their alloys.

Method of annealing large metal blocks by electric induction heating. (C. LORENZ A-G. Wilhelm Fischer, inventor). Ger. 515,913, July 3, 1928.

10-ORGANIC CHEMISTRY

CHAS. A. RUTHER and CLARENCE J. WEST

Methane synthesis from carbon monoxide and water vapor. P. PASCAL and E. BOTTIGER. J. chim. phys. 54, 457-8 (1957). CO and H_2O give quant. yields of CH_4 and CO_2 in the presence of NiO at 275° and practical yields up to 75°. At higher temps. equal molar mixts. of CH_4 and H_2 are formed, until at 750° H_2 is quantitatively obtained.

Properties of conjugated compounds. II. Addition of hydrogen bromide to β,γ - and α,β -dimethylbutadiene. ERNEST H. FARMER and FREDERICK C. B. MARSHALL. J. Chem. Soc. 1931, 123-37. Cl. C.A. 24, 724. The mode of addn. of HBr to β,γ -dimethylbutadiene (I) has been studied by BERGMANN (C.A. 17, 142) and by CLAUSSEN (C.A. 17, 1419). B. states that I yields a mixture of α - and β -isomers, $\text{Me}_2\text{C}(\text{Br})\text{CH}(\text{Me})\text{CH}_3$ (II), with 1 mol. HBr, while C. concludes that, although an unstable II is immediately formed, the actual product is the α -isomer, $\text{Me}-\text{C}(\text{CH}_3)=\text{CH}-\text{CH}_2\text{Br}$ (III). The chem. evidence obtained in the study by the degradation of a no. of distd. and undistd. specimens of the acids, product of HBr to I points unmistakably to the fact that the primary formed product (even when addn. is effected at -27°) in the absence of excess of HBr is III, it also appears to be the sole product, although the complete absence of very small proportions of II could not be asserted. The α -isomer gives a mixture of $\text{Me}-\text{CH}(\text{Br})\text{CH}(\text{Br})\text{CH}(\text{Me})\text{CH}_3$ (IV) and $\text{Me}-\text{CH}_2\text{CH}(\text{Br})\text{CH}(\text{Br})\text{CH}_3$ (V), as shown by conversion into the glycols and mono- and diacids. The 3rd isomer was absent, showing that α,β -addn. takes place in only 1 of 2 possible ways. The ratio of IV to V in the reaction product, as shown by the ratio of AcOH to EtCO_2H formed, is approx.

00 10, this figure agrees roughly with the highest estimate of the proportion of AcII produced by ozonolysis of the hydrobromide. It is clear that the manner in which the alkylbutadienes add HBr varies, as it does for other sym and unsym alkenes, with the position of substituent groups in the unsat chain. At present there is nothing to indicate the order of events in the 2 fold process of addendum attachment and addendum partition.

C. J. WENT

Mercaptan chemistry. WILLIAM M. MALISOFF, ERNEST M. MARKS and IRID G. HESS. *Chem. Reviews* 7, 493-547 (1930).—A crit presentation and discussion of the SH compds under the following headings: (1) phys characteristics, (2) chem characteristics, (3) methods of prepn and purification, (4) detection and detn, (5) special interest attaching to the substances, (6) research problem. Approx 600 references are given.

G. TOINNIER

Nitrosites and nitrosates. LYDIA MONTE. *Gazz. chim. Ital.* 60, 787-97 (1930).—The present paper deals with the prepn of new nitrosates and nitrosites of octylene (I), disobutylene (II) and hexadecylene (III), the formation of addn products between these hydrocarbons and N_2O , and the decompn of the nitrosates and nitrosites by heating in an inert gas. The literature describes only the nitrosates and nitrosites of the lowest members of the alkyl hydrocarbons up to amylene (cf. Vegorov, *C. A.* 7, 1477; Wallach, *C. A.* 9, 1158). Following the usual method of prepn, the final products from I, II and III are yellow oils with the odor of fatty acids, and in certain cases the acids were identified. This formation of acid is attributed to the direct oxidizing action of N_2O or to decompn with autooxidation which nitrosates and nitrosites (formed at first) undergo spontaneously at room temp (cf. Schaarschmidt and Hoffmeier, *C. A.* 19, 2631). The expts prove that addn products are formed. The gases evolved in the decompn of the nitrosates and nitrosites contained in all cases N , NO and CO_2 and no other gases in any appreciable quantity (cf. Schaarschmidt and Hoffmeier, *loc. cit.*; Sommer, *Ber.* 29, 357 (1896)). The formation of N and CO_2 indicates autooxidation and destruction of the mol. The quant results are of special interest, for though the components are always the same, the proportions of the gases vary greatly. One of the chief aims in the investigation is to develop a reliable method for distinguishing nitrosates from nitrosites. I was prepd by refluxing *sec*-octyl iodide (100 g) with alc. $NaOH$ (25 g in 500 cc) for 10-2 hrs, pouring into a large vol of water, drying the sepd oil with $CaCl_2$, rectifying and recovering the 121-3° fraction. II was prepd by the method of Butlerow (*Ann.* 182, 44 (1877)) and III by the method of Bertholet and Péan. In numerous expts nitrous vapors (from As_2O_3 and HNO_3 of d 1.38) were bubbled through solns of III in Et_2O or $AcOH$ (5-10 parts of solvent). The liquid became green, but only under certain conditions in $AcOH$ was a cryst compd, m 60-7°, obtained in too small quantity to be analyzed. When the solvents were evapd. at room temp, thick yellow green oils were obtained, and these dissolved in Et_2O , washed with 5-6% aq Na_2CO_3 , the alk ext acidified with dil HCl , extd with Et_2O , the ext evapd, the residual oil crystd by cooling, dissolved in 95% $EtOH$, neutralized with alc. $NaOH$ (phenolphthalein), alc. $AgNO_3$ added, warmed, filtered, the residue washed with $LiOH$, dissolved in NH_4OH , evapd, the ppt washed and dried *in vacuo*, yields *Ag pentadecylate*. When the crude oil is let stand several months (CO_2 II), is deposited (Schaarschmidt has obtained (CO_2 II), by the slow action of N_2O on satd aliphatic hydrocarbons, cf. *Z. anorg. Chem.* 37, 918 (1924), cf. *C. A.* 18, 2138). Similar results were obtained by treating cold liquid N_2O_4 , by attempting to prep the addn product with $AmNO_3$ and HNO_3 , by treating III in an inert solvent with $NaNO_3$ and H_2SO_4 , and by absorbing N_2O in III under slight pressure at room temp. I and II also gave similar results. The oil from I contained fatty acids, but no Ag salt in a pure enough form to be analyzed was obtained. Here, too, (CO_2 II), was formed on long standing. With II a thick blue liquid was obtained. Thus, dil H_2SO_4 (7 cc concd H_2SO_4 in 40 cc water) added to II (8 g) in ligroin (40 cc), made ice-cold, aq $NaNO_3$ (10 g in 20 cc) added dropwise, the aq part removed, the ligroin soln evapd. *in vacuo* over CaO and paraffin, yields a blue liquid which after 8-10 days becomes a greenish yellow thick oil. Treating this oil in the same way as that from III, a Ag salt contg 59.8% Ag was obtained, probably a mixt of salts. Similar blue liquids were obtained from I and III. The d of these products is over 1, they do not react with water and they give the Liebermann reaction. They are also formed by bubbling nitrous vapors (from As_2O_3 and HNO_3 of d 1.38) into ice cold I, II or III without solvent, stopping the reaction when the liquid has become dark blue, washing the product with water and sepg by difference in d. The stability of the products varies with the nature of the hydrocarbon. That from II is the most stable, and the color lasts for several days. That from III is much less stable and the color disappears

after a few hrs. The mol. wt. of the product from II was 179.2, which agrees fairly well with that for the nitrosite $C_{10}H_{15}O_2N$. The formation of addn products was also proved by detn. of the quantity of N_2O_4 by II at room temp. Quant. data on the absorption of N_2O_4 by I, II and III at room temp. show in all 3 cases absorption which corresponds fairly closely with the addn. of 1 mol. of hydrocarbon with 1 mol. of N_2O_4 . That true addn. products are formed is proved by mol.-wt. detns. Thus the mol. wt. of the product from III was 304 (316 calcd. for $C_{10}H_{15}N_2O_4$). A study was then made of the decompn. of nitrosites and nitrosates, for which purpose amylene nitrosate (IV) and the isomers styrene α nitrosite (V) and styrene β nitrosite (VI) were chosen, for these can be obtained in a very pure form. Further expts. will deal with other analogous compds. IV was prepd. by the method of Wallach (cf. *Ann* 241, 292(1887)) and V and VI by the methods of Wieland (cf. *Ber* 36, 2359(1903)). The app. for their decompn., which is described in detail and illustrated, allows heating of the nitrosate or nitrosite in the dry state in a current of inert gas (CO_2 or N_2). IV was thus heated at 90-5°, V at 100-25° and VI at 125-35°. In all 3 cases, N_2 , NO and H_2O were evolved, and in no case was any appreciable quantity of a higher oxide of N , of O or of a combustible gas detected. CO_2 was evolved in considerable quantity. The following data give the % N_2 (by wt.) evolved, the % NO evolved, the % total N evolved and the ratio % N evolved/% N in compd., resp. for IV, V and VI: IV, 5.84, 2.50, 7.00, 0.405, V, 3.28, 6.93, 6.61, 0.418, VI, 4.04, 4.85, 6.30, 0.405. These results show that, although the components in the gaseous decompn. products are in all cases the same, the % compn. varies with the starting substance. In each case, the % N evolved in gaseous form (N_2 + NO) is around 40% of the total N in IV, V or VI. The remaining 60% of solid decompn. products is composed of NO_2 compds., nitriles, etc.

C C DAVIS

Attempted preparation of methanetetrasulfonic acid. FREDERIC B. KIPPING. *J. Chem. Soc.* 1931, 222-3.— $I_2C(SO_3K)_2$ and K_2SO_4 in H_2O heated 4-5 hrs. give K methionate. $AcNH$, and H_2SO_4 (65 or 30% SO_3) at various temps. give $CH(SO_3K)_3$, with 10% SO_3 , below 170-80° only $AcOK$ could be isolated but at 210° K methionate was obtained in a poor yield.

C J WEST

Changes in properties of chemical compounds by complex formation. VI. Methylation of alcohols by diazomethane. HANS MEERWEIN and GERHARD HINE. *Ann* 484, 1-25(1931), cf. C A 23, 4189.—Aliphatically bound HO groups, as far as is known, do not react with CH_3N_2 (I). The following results were obtained with 0.5% EtOH I at 0° (cc. N_2 in 1 min.): EtOH, 0.3, PrOH, 0.1, iso-PrOH, 0.18, BuOH, 0.18. Sufficient BuOH does not react with I in 30 hrs. to make the method useful for the prepn. of the ether. Since it would appear that an activation of the HO group might be brought about by complex formation or by the introduction of a polar or easily polarized group in the neighborhood of the HO group, a study of the latter possibility was undertaken. The following results indicate that this is the case and that the alcs. studied were more or less easily methylated by I. In general, the reaction was carried out by passing the I directly into the alc. Ethers retard the methylation in all cases and for this reason the reaction does not go to completion, the resulting ether preventing further methylation. $PhCH_2OH$ gives about 13% of the Me ether $(CH_3OH)_2$, and its mono-Me ether scarcely reacts with I. $HIOCH_2CH_2OAc$ in EtOH gives a mixt. of $AcOEt$, $HOCH_2CH_2OMe$, $MeOCH_2CH_2OAc$ and $(CH_3OH)_2$. Without a solvent the yield of the Me ether was about 12%. $AcOMe$ (?) was also formed, although no explanation of its formation can be given. α -Methoxyacetone gives 25% of the mono Me ether, b_p 126-9°, and some di Me ether b_p 100-10° ($CH_3)_2O$ (30 g.) and 42 g. Cl_2CCO_2H give 15 g. of trichloroacetyl glycol (II), b_p 130-40° with I 12 g. II gives 5 g. of the Me ether $Cl_2CC(OCH_3)_2OMe$, in 77.8°, which is not changed by boiling H_2O or N NaOH but splits off Cl_2CCO_2H with 0.25 N HNO_3 . $ClCH_2CH_2OH$ (15 g.) gives 8 g. of the Me ether in 91°. $ClCH_2CH(OH)CH_2OH$ (30 g.) gives 2.5 g. of the Me ether, b_p 171-2°, d_4^{20} 1.1832, and 2.3 g. of the di Me ether, with twice the amt. of I there results 2 and 1 g. resp. $(CH_3)_2CHOH$ (30 g.) gives 26 g. of the Me ether, b_p 159.5-61.5°. $HOCH_2CH_2CN$ and I react vigorously, giving in addn. to the Me ether (III) b_p 162-4°, basic compds. with a C_4H_5N odor. MeO detns. indicate a 52% yield of III, although but 25% could be obtained pure. There was also isolated a compd., C_4H_5ON , as the HCl salt. $MeCH(OH)CN$ gives about 20% of the Me ether, b_p 113-5° d_4^{20} 0.8942, n_D^{20} 1.3825, Me_2CO , $MeEtCO$ and $Me_2CO \cdot CH_3$. $AcCH_2OH$

gives only a small amt. of the Me ether, the principal product (50-60%) being $EtCOCH_2OH$. $BzCH_2OH$ behaves in a similar manner, giving phenylacetylcarbinol, b_p

140-5°. With 20 g $\text{HOCH}_2\text{CH}_2\text{Cl}$ as a catalyst, 20 g PrOH and 8 g I give 5 g PrOMe and 5.6 g $\text{MeOCH}_2\text{CH}_2\text{Cl}$. BuOH (20 g), 10 g $(\text{ClCH}_2)_2\text{CHOH}$ and 7 g I give 3.2 g BuOMe and 7 g $(\text{ClCH}_2)_2\text{CHOMe}$. BuOH (10 g), 7 g $\text{ClCH}_2\text{CH}_2\text{OH}$ and 5 g I give 2.3 g BuOMe and 4.8 g $\text{ClCH}_2\text{CH}_2\text{OMe}$. Ten g of a 5% BuOH ZnCl_2 soln. and 2.2 g I give 2.0 g BuOMe . 20 g of a 3.8% soln of FeCl_3 in BuOH and 7 g I give 6.1 g BuOMe . 20 g of a soln of MgCl_2 in BuOH give 5 g BuOMe . A 5% soln of Al(OLt) , in BuOH and I give 83% of BuOMe . iso-PrOH gives 77% iso-PrOMe . B(OLt) , in tOH (15 g of 5% soln) and I give 13 g tOMe , 0.69 g B(OMe) , and 13.8 g PrOH with 1 g I give 4.1 g PrOMe . iso-PrOH and BuOH are not methylated in the presence of B(OLt) . Sb(OLu) and BuOH give 73% of BuOMe . iso-PrOH scarcely reacts with I in the presence of iso-PrOSb . C. J. WEST

Some aliphatic compounds of arsenic. W. J. CROCH DYE, GWYN DAVIES AND W. J. JONES. *J. Chem. Soc.* 1931, 1438— Pr_3AsH_2 , $d_4^{25} 1.4192$, $b_{25} 113^\circ$, which consists, are related by the equation $10 \log p = 3215.7 + 273$. Exposed to the air, the arsine gradually deposits the oxide, also prepd by shaking the arsine with H_2O or the dibromide with Ag_2O . *Tributylarsine oxide*, pearly crystals, and the *Am* deriv were thus prepd. *Tripropyl* and *triisobutylarsine sulfides* crystallize as long needles. Pr_3AsBr_2 and Ag_2SO_4 in tOH give *triisobutylarsine sulfate*, $m 84^\circ$. *Dichlorides* result from the arsine and Cl in CCl_4 , *tri-Pr* deriv, $m 81^\circ$, *tri-Bu* deriv, $m 40^\circ$, *tri-iso-Bu* deriv, $m 130^\circ$, *tri-Im* deriv, sweet smelling liquid. The *dibromides* prepd similarly, are highly hygroscopic. Pr_3AsBr_2 is expelled by Cl , HNO_3 or H_2SO_4 , *tri-Pr* deriv, $m 95^\circ$, *tri-Bu* deriv, $m 55^\circ$, *tri-iso-Bu* deriv, $m 135^\circ$. *tri-Am* deriv, liquid. The *disulfides* are formed in light petroleum; they become brown on exposure to the air, *tri-Pr* deriv, m about 130° , *tri-Bu* deriv, $m 121^\circ$, *tri-iso-Bu* deriv, $m 117-0^\circ$. HgCl_2 gives compds of the general formula $\text{R}_3\text{As}_2\text{HgCl}_2$, which may be recrystd from EtOH ; *tri-Pr* deriv, $m 106^\circ$, *tri-Bu* deriv, $m 70^\circ$, *tri-iso-Bu* deriv, $m 158^\circ$. The arsines and MeI in Et_2O give *methoxides*, *methyltriethylarsonium iodide*, $m 270^\circ$, *tri-Pr* deriv, $m 204^\circ$, *tri-iso-Bu* deriv, $m 254^\circ$, in less pure condition were prepd the *tri-Bu* and *tri-Am* derivs, also *ethyltriethylarsonium iodide*, m about 107° . *Methyltriethylarsonium chloroplatinate*, $m 219^\circ$ (decompn), *tri-iso-Bu* deriv, orange, $m 224^\circ$ (decompn), *tri-Am* deriv, orange, $m 114^\circ$. The arsine *methoxides* and CdI_2 give the *arsonium cadmi-iodides*, best crystal from EtOH , *methyltriethylarsonium cadmi-iodide*, 2MeEtAsI CdI_2 , $m 23^\circ$, *tri-Pr* deriv, $m 279^\circ$, *tri-Bu* deriv, $m 100^\circ$, *tri-iso-Bu* deriv, $m 138^\circ$. *Methyltriethylarsonium mercurichloride*, $\text{MePr}_3\text{AsCl HgCl}_2$, $m 141^\circ$. *Methyltriethylarsonium mercuriodide*, $m 61^\circ$, *tri-Pr* deriv, $m 90^\circ$, *tri-Bu* deriv, $m 76^\circ$, *tri-iso-Bu* deriv, $m 117^\circ$. C. J. WEST

Parachor and chemical constitution. XVI. Silicon compounds. SAMUEL SUGDEN AND HENRY WILKINS. *J. Chem. Soc.* 1931, 126-8, cf *C. A.* 24, 2651— Si(OMe)_4 , $b_{25} 121-2^\circ$ (all b p cor.), $d_4^{25} 1.032$, $\gamma^{25} 24.98$, parachor (P) 330.9, $d_4^{25} = 1.067 - 0.00155t$, Si(OEt)_4 , $b_{25} 165.3-5.8^\circ$, $d_4^{25} 0.933$, $d_4^{25} = 0.957 - 0.00135t$, $\gamma^{25} 22.21$, $P 497.6$. SiEt_4 , for which a method of prepn is given, $b_{25} 153^\circ$, $d_4^{25} 0.7563$, $d_4^{25} = 0.7818 - 0.000791t$, $\gamma^{25} 23.69$, $P 412.1$. SiPr_4 , $b_{25} 213-5^\circ$, $d_4^{25} 0.7809$, $d_4^{25} 0.7090 - 0.000727t$, $\gamma^{25} 24.65$, $P 565.3$. SiPh_4 , $m 231^\circ$, $d_4^{25} 0.937$, $d_4^{25} 0.983 - 0.000883(t - 200)$, $\gamma^{25} 23.70$, $P 787.5$. The mean atomic P for Si is 27.8. Various values are given for d and γ for these compds. C. J. WEST

Preparation of fatty acids. R. LUKES. *Chem. Listy* 24, 197-200 (1930)—A general method of synthesis of fatty acids from alcs contg 4 C atoms less than the desired acid is as follows. The alc is converted into the alkyl bromide, and this is combined with Mg to yield the corresponding Grignard reagent, which reacts with *N*-methylsuccinimide to yield a 1-methyl-5-alkylpyrrol-2-one. This on hydrolysis yields a γ -keto acid possessing 4 C atoms more than the original alc., and this on electrolytic reduction yields the corresponding fatty acid. Hexoic, heptoic, octoic, nonoic and decaoic acids have been prepd in this way. The following hitherto unknown intermediate products are described: 1-methyl-5-amylpyrrol-2-one, $b_{25} 143-8^\circ$, 1-methyl-5-hexylpyrrol-2-one, $b_{25} 148-50^\circ$; γ -ketononoic acid, $m 69^\circ$, and γ -ketodecoic acid, $m 70-1^\circ$. B. C. A.

The mechanism of the decomposition of alkaline chlorohydrins. LENNART SMITH. *Z. physik. Chem., Abt. A*, 152, 153-6 (1931)—Doubt that this reaction is not monomol seems not to have been well founded, for S , by using pure samples, has obtained good constants for $(\text{CH}_3\text{OH})_2$ and $\text{Me}(\text{CH}_2\text{OH})_2$. V. F. HARRINGTON

Formins of glycerol. P. A. DUBOIS. *Thesis*, Paris, 1929, *J. pharm. chim.* [8], 12, 478-9 (1930)—The objective was a rational and practical method of prepg. *allyl alc*. (1) Summary: (1) The mechanism of the formation of I , starting from HCO_2H and glycerol, is different from that when $(\text{CO}_2\text{H})_2$ is used. (2) The esterification of

glycerol by HCO_2H leads principally to the formation of diesters. (3) HCO_2H seems to act more easily on the secondary alc function of glycerol since the esters which predominate among the isomers of the same degree in the reaction mixt, are 2 mono formun and 1,2-diformun. (4) When esterification is pushed to the diformun stage, a yield of 70% I is easily obtained.

A new form of the two antipodal rhamnitol. F. VALENTIN. *Collection Czechoslov. Chem. Comm.* 2, 689-95 (1930).—Hydrated sugar alcs having definite compn and enantiomorphous cryst forms have been prep'd for the first time. I Rhamnitol formed by reduction of ordinary rhamnose with 3% Na II and crystd slowly from dil H_2O soln forms the trihydrate (I), m 69° , $[\alpha]_D$ 8.83° . The H_2O of crystn was det'd by drying at 100° in *vacuo*. Crystals of *d* rhamnitol trihydrate (II), m 69° , $[\alpha]_D$ -8.89° , were obtained similarly from *d* rhamnose (see C A 25, 81). I is rhombic, a b c , 0.9947 1.11750. The measured angle between the optic axes for Na light is 82.75° . The indices of refraction are α_D (calc'd) 1.44, β_D 1.492, γ_D 1.505. The axial ratios and optical properties of II are the same.

JANET E. AUSTIN.
Dipentaerythritol. WALTER FRIEDRICH AND WILHELM BAUM. *Ber* 63B, 2681-90 (1930).—According to Ger pat 390,622, pentacerythritol (I) is prep'd by adding to AcH and 4 mols HCHO in dil aq soln at 15° 1 equiv alkali or alk earth hydroxide in the course of 1 hr and slowly raising the temp to 45° during the next hr. The m p of the crude product obtained by evapn the reaction mixt is raised to about 210° by crystn from H_2O but again falls on further crystn. It was concluded that I is accompanied by an impurity which is less sol in H_2O . This by-product has now been shown to be the ether, dipentaerythritol, $\text{O}(\text{CH}_2\text{C}(\text{CH}_2\text{OH})_2)_4$ (II). Since the I and II cannot be sep'd by crystn, the crude product was nitrated and the nitrates sep'd through their different soly in Me_2CO . Treatment of the nitration product with Me_2CO leaves undissolved almost pure I tetranitrate (III), the last traces of which are ppt'd by cautious addn of H_2O to the Me_2CO soln. This produces 2 layers: a H_2O - Me_2CO mixt contg a little II hexanitrate (IV), and a sat'd soln of almost pure IV in Me_2CO and a little H_2O . The latter stirred into 8-10 vols alc yields first a yellowish brown sirup and then white crystals, and the sirup, reprec'd from Me_2CO with alc, gives a pure white product which is recrystd from alc in *vacuo*. The IV cannot be saponified with alc, KOH because of the oxidizing action of the resulting KNO_3 on the II, but hydrolysis was effected with Zn and HCl or, better, with cryst Na_2S in boiling Me_2CO - EtOH . The II (yield, 89%), m 221° , mol wt in boiling H_2O 250 d_{20} 1.306 soly in 100 parts H_2O 0.12 and 11.64 at 15° and 100° , in alc 0.007 and 0.31 at 20° and the b p, resp. Perfectly pure I, whose prep'n is to be described elsewhere, cannot be obtained by recrystn, a prep'n which no longer gives the fuchsian SO_4 test for II (see below) m 210° . The m-p curve of mixts with II falls steeply to a min at about 190° with 30% II and shows that an I m 235° , which for practical purposes may be considered pure, contains 10% II. Radulescu's statement that the by-product in 220° obtained in the prep'n of I can be converted into pure I, m 251° , by treatment with alc was not confirmed, the m p of pure II was not changed in the slightest by several crystns from alc. IV (95.7% from II slowly added to cold HNO_3 (d 1.52) and treated with $\frac{1}{2}$ vol conc'd H_2SO_4), m 75° , d_{20} 1.630 (cryst) 1.017 (fused), mol wt in freezing C_2H_5 522, the f-p curve of mixts with III (m 141°) is almost a straight line with no min or max. Hexaformate, from II refluxed in conc'd HCO_2H and fumed down twice with HCO_2H at 120° , yellow m 56° . Hexaacetate, from II, Ac_2O and NaOAc , heated under a reflux until a violent reaction sets in (yield 54.8%), m 73° , mol wt in C_2H_5 514. Hexabenzooate, from II, BzCl and NaOH below 40° (83.5% yield), m 183° . Hexatrisyl ether (97.3% from II and Ph_3CCl in $\text{C}_2\text{H}_5\text{N}$ at 100°), m 173° , decomp'd by atm moisture and completely hydrolyzed by boiling H_2O . Attempts to increase the yield of II by the use of other alk condensing agents (KOH, $\text{Ba}(\text{OH})_2$, $\text{Pb}(\text{OH})_2$, $\text{Zn}(\text{OH})_2$, Na_2CO_3) were unsuccessful, boiling for days with NaOH, $\text{Ca}(\text{OH})_2$, HCO_2Na , $(\text{HCO}_2)_2\text{Ca}$ did not change the m p of pure I; fusion of I with K_2CO_3 yielded a brown resin contg no II or other cryst product. No II was formed when I was heated at 300° under a steam jacketed head which permitted the H_2O to distil off and condensed the higher boiling vapors back into the flask, nor when the I was heated at 300° in a sealed tube with 0.25 part H_2O . Pentaerythrose, condensed with AcH and HCHO in the most varied proportions, gave only ordinary crude I and in poorer yield than the method described above. The mixt of glyceric aldehyde and $\text{CO}(\text{CH}_2\text{OH})_2$ formed by oxidation of glycerol yielded no cryst products. $\text{MeCH}_2\text{CH}_2\text{CHO}$ with 8 mols HCHO and 1 equiv CaO gave I, m $218-52^\circ$, in 64% yield as det'd by conversion into III. CH_2 , CH_2CHO with 3 mols HCHO and 0.5 mol $\text{Ca}(\text{OH})_2$ gave 63.4% I. Expts with AcH , 0.5 mol $\text{Ca}(\text{OH})_2$ and varying amts (1-8 mols)

NIIOH, and why only aq. III attacks Ni, Co, Cu and Fe (cf. Note VI, *C. A.* 17, 2268), or, in other words, why I has a basic function in which an oximino N atom becomes quinquevalent and III an acid function toward metals of the 8th group. The method of prepn of I and III already described (cf. Note VI, *C. A.* 17, 2268) should be modified as follows. Aq. HONH₂Cl (5.3 g.) and aq. AcONa (11.8 g. in 15 cc.), added to hot alc. IV (10 g. in 20 cc.), kept at 60–80° for 12 hrs. (μ_n is 5.2–5.4), cooled, the ppt. (0.2 g. of a mixt. of 95% I and 5% III) dissolved in anhyd. Et₂O, made ice-cold, dry HCl passed through the soln., and the ppt. treated with dil. Na₂CO₃, yield 6 g. of pure I. The mother liquor, concentrated to the point of crystn., yields on cooling 4 g. of a mixt. of 75% of III and 25% of I. The mixt. dissolved in boiling 50% EtOH, glacial AcOH added, excess hot 21% Ni(OAc)₂ added, the ppt. washed with EtOH, decompd. with HCl and extd. with Et₂O, yields 2.8 g. of III. To isolate the I which remains in the AcOH soln. from which the Ni complex of III seps, the soln. is made basic with NH₄OH, the ppt. is treated with dil. H₂SO₄ and extd. with Et₂O. This yields 1.2 g. of slightly colored I. Avogadro (cf. Notes XIII and XLII, *C. A.* 18, 661, 21, 1976) has prepd. α - and β -tolylglyoximes (V and VI, resp.). Dry HCl passed into V in ice-cold anhyd. Et₂O, does not ppt. a HCl salt, but 40% of V isomerizes to VI, and the latter is readily isolated as the Ni complex, Ni(C₆H₄O₂N₂)₂. Following the same procedure with VI, 60% of VI isomerizes to V. To isolate the V, the product is washed with Na₂CO₃, the solvent eliminated, the residue dissolved in boiling 50% EtOH, a little glacial AcOH added, excess Ni(OAc)₂ then added, the ppt. (Ni complex of VI) removed, the filtrate made alk. with NH₄OH, the ppt. treated with dil. H₂SO₄ and finally extd. with Et₂O. V apparently m. 170–1°, but it has no true m. p. since by the action of heat it isomerizes to VI. LXX. *Ibid.* 895–93.—A better method for the prepn of PhC(NOH)C(NOH)Cl (I) than that previously described (Note VI, *C. A.* 17, 2268, Note XXIV, *C. A.* 20, 746) is to dissolve α - or β -phenylglyoxime (or their mixt. from the oximation of PhCOC(NOH) in a min. of 94–9% AcOH, cool, pass into the soln. a slight excess of Cl₂, wash the ppt. with AcOH and dry at 70–80°. It m. 190–200°. I and BzCl (calcd. quantity), heated 1 hr. on a boiling water bath and the product recrystd. from CHCl₃, yields 2 benzoylphenylchloroglyoxime, PhC(NOBz)C(NOH)Cl (II), m. 177–8° (slight decompn.), does not form an Ac deriv. with Ac₂O at room temp. II and excess BzCl, heated for a long time at 100° or boiled for a few min., brown added and the ppt. recrystd. from EtOH, yield *dis*benzoylphenylchloroglyoxime (III), m. 148–9°. III, dissolved in warm alc. NH₃, the solvent eliminated, the residue dissolved in dil. HCl, cryst. AcONa added and the ppt. recrystd. from water, yields β -phenylaminoglyoxime (IV) (cf. Note VI, *C. A.* 17, 2268). II in Et₂O, agitated with a slight excess of 6 N NH₄OH for a short time, and the ppt. recrystd. from EtOH, yields 2 benzoyl α -phenylaminoglyoxime, PhC(NOBz)C(NOH)NH₂ (V), m. 171–2° (slight decompn.), insol. in cold dil. NaOH, in EtOH with FeCl₃ it gives a light violet brown color which disappears when dil. with water. V, heated gently with 20% NaOH and a little EtOH, dil. with water and the ppt. recrystd. from water or very dil. EtOH, yields phenylaminofurazan (cf. Note X, *C. A.* 17, 3875). The alk. mother liquor, neutralized with CO₂, extd. with Et₂O the ext. evapd. and the residue recrystd. from a mixt. of Et₂O and petr. ether, yields α -phenylaminoglyoxime (VI) (cf. Note VI, *C. A.* 17, 2268). Alc. II and a slight excess of PhNH₂, heated to boiling dil. with water, acidified with dil. HCl, filtered (the residue is BzNH₂). NaOAc added to the filtrate and the ppt. recrystd. from CHCl₃, yield α -phenylaminophenylglyoxime (cf. Note VIII, *C. A.* 18, 661). In the cold, PhNH₂ does not react with II. Cold IV, Ac₂O and fused AcONa yield, after crystn. of the product from EtOH, 2 benzoyl *N*-acetyl α -phenylaminoglyoxime, PhC(NOBz)C(NOH)NHAc (VII), m. 190–1° (slight decompn.). That the Ac group is united with the amino N and not to the oximino O is shown by the fact that when VII is heated gently with 20% NaOH and a little EtOH phenylaminofurazan and its Ac deriv. are obtained. In conjunction with earlier expts. (cf. Notes XXXVII, LXIV and LXV, *C. A.* 21, 1099–25, 79–80), the present expts. show that when BzCl acts at 100° on chloroglyoximes RC(NOH)C(NOH)Cl, where R is H, Me or Ph, the oximino H farthest from the Cl is first replaced by Bz, giving the RC(NOBz)C(NOH)Cl with which in turn excess BzCl forms RC(NOBz)C(NOBz)Cl. The constitution of V corresponds to VI, i. e., it is a deriv. of a form of VI which in the free state is to be regarded as a labile form which is easily isomerized by fusion or by heating in dil. AcOH into IV, a stable form. Since I is transformed by NH₃ into IV, which is not isomerized by NH₃ to VI, in the benzoylation at 100° of I a 2 Bz deriv. is obtained of a form of I incapable of existing in the free state. This bears the same relation to the only known form of I as does VI to IV. HCl can isomerize 1 form [of glyoxime to the other form (cf. Note LXIX, preceding abstract), so it is not impossible that a

identical behavior, agrees with the existence in furoxans of both ter- and quinquevalent N and in dioxidiazines of 2 trivalent N atoms, and shows that the $C_2N_2O_2$ group of furoxans has an asym structure and that of dioxidiazines a sym structure. Besides confirming the structures of furoxans and dioxidiazines shown above, *s e.*, of the isomeric methylaryl peroxides, the expts exclude the possibility that the peroxide

with the lower *m p* can be the furozan $MeC \overline{N} O N(O) CAr$. The latter would

behave like the other furazan $PhC \overline{N}(O) N CAr$, and yield with the Grignard reagent MeCN, which in numerous expts could not be obtained, either as such or as its hydrolysis product AcOH, by the use of MeMgI, EtMgI or PhMgBr. The reaction between methylaryl peroxides and the Grignard reagent is difficult to interpret. If MeMgI attaches to an ethylenic double bond and thus brings about its rupture, the 2 isomeric peroxides have the same structure



like that proposed by Green and Rowe for *o* quinone dioxime peroxides (cf *C A* 7, 3119, 8, 1412). This structure may, however, be that of the peroxides with the lower

m ps which have been regarded as dioxidiazines $MeC \overline{N} O O N CAr$. With triphenylisoxazoline oxide, MeMgI adds to the $-C=N=O$ group (cf Green and

Rowe, *loc cit*) and then with water forms a compd in which the bond between C and N still exists. The same is true of the methylaryl peroxides with the high *m ps* (furoxans), but is not true of those with the lower *m ps*. The latter do not, therefore, have this structure, *s e.*, are not furoxans. *Exptl*—Methylphenylfurozan (5.5 g.), added cautiously to MeMgI in Et_2O (from 21 g MeI, 250 g Et_2O (distd over Na) and 3.4 g Mg), let stand after the reaction has subsided, the Et_2O evapd, ice water added, steam-distd and the distillate collected in dil HCl, yields insol PhCN. The latter is removed, and the mother liquor, boiled, yields AcMe in the distillate while the dry residue is NH_4Cl . Under the same conditions methyl *p*-bromophenylfurozan yields *p*-BrC₆H₄CN, AcMe and NH_3 , and 4 methyl 5-*p*-bromophenyl 1,2,3,6-dioxidiazine (III) yields *p*-BrC₆H₄Ac, AcMe and NH_3 . Crystd from hot $EtOH$, III *m* 92° (cf 89–9° of earlier expts, Note L, *C A* 23, 375). Similarly, methyl *p*-methoxyphenylfurozan (IV) and MeMgI yield *p*-MeOC₆H₄CN (V), AcMe and NH_3 . With EtMgI IV reacts only with difficulty, with formation of V, MeCOEt and NH_3 . 4 Methyl 5-*p*-methoxyphenyl 1,2,3,6-dioxidiazine (VI) and MeMgI form *p*-MeOC₆H₄Ac, AcMe and NH_3 . With EtMgI, VI reacts only with difficulty, with formation of *p*-MeOC₆H₄COEt, MeCOEt and NH_3 . Methyl[3,4-dihydroxyphenyl]furozan methylene ether and MeMgI yield piperonyl acetone, AcMe and NH_3 . 4 Methyl 3-[3,4-dihydroxyphenyl] 1,2,3,6-dioxidiazine methylene ether and MeMgI yield 3,4-CH₂O₂C₆H₃COMe (acetopiperone), AcMe and NH_3 . LXXII. E. DUBO AND M. BISSI. *Ind* 899–903.—In the oxidation of *o*-benzil dioxime (I) with alk $K_2Fe(CN)_6$, Auwers and Meyer (cf *Ber* 21, 806 (1888)) obtained the diphenyl peroxide $Ph(C_2N_2O_2)Ph$ (II) and a secondary product $C_{12}H_8O_2N_4$ (III). III has never been described since, yet its formation may be of importance in explaining the dehydrogenation of glyoximes to the so-called peroxides. The present paper shows that III is formed from the α but not from the β -form of I, and describes the conditions which lead to a much higher yield than the extremely small yield by the procedure of Auwers and Meyer. Aq 15% $K_2Fe(CN)_6$ (30 g) added dropwise to I (10 g) in 10% KOH (200 cc) keeping ice-cold, the ppt. purified by boiling in $EtOH$ and recrystg the residue from glacial AcOH, yields 25–30% of the compd $C_{12}H_8O_2N_4$ 2AcOH in air it loses slowly at room temp and rapidly at 100° its AcOH of crystn, leaving III, which was proved to be dioxotetraphenylpyrazine

$O \overline{N} CPh CPh N(O) CPh CPh$. It is yellow, *m* 322°, gives an intense red soln in concd H_2SO_4 (from which it is reprecipd by water), is not altered by heating with HCl (d. I 19) at 160–70° in a sealed tube. Traces of III are also formed by the oxidation of I with NaClO (cf Note LXI, *C A* 24, 3488). III (5 g) in glacial AcOH and Zn dust (3 g), heated several min on a boiling water bath, filtered and the filtrate cooled, ppts 3.5 g of tetraphenylpyrazine (IV). After sepn of IV, the mother liquor, made alk

with NaOH and steam distd., yields a small quantity of tetraphenylpiperazine III and PCl_5 (equal parts), heated at $140-50^\circ$, yield a brown liquid which treated with water solidifies, and crystd from glacial AcOH, yields chlorotetraphenylpyrazine (V), m 212° gives an intense red soln in concd H_2SO_4 . This reaction is similar to that of furoxans under the same conditions (cf. Notes L and LVI, C. A. 23, 375, 24, 815), i. e., III is first deoxygenated to IV, which then reacts with Cl with formation of V. The formation simultaneously of II and III from I proves that $\text{K}_2\text{Fe}(\text{CN})_6$ acts on I in 2 distinct ways (1) simple dehydrogenation which forms II, and (2) a more complex reaction which involves the elimination as HNO_2 of 2 H atoms and 2 NOH groups from 2 mols of I, thus $2\text{I} + 2\text{O} \rightarrow \text{III} + 2\text{HNO}_2 + \text{H}_2\text{O}$. Since under the same conditions β and α benzil dioxime (VI and VII) form exclusively II, both NOH groups of VI and VII are dehydrogenated by oxidizing agents, whereas in I, 1 NOH group is dehydrogenated and the other is oxidized, i. e., toward oxidizing agents the NOH groups of VI and VII behave the same and those of I differently. This behavior is analogous to that with N^{++} ions where of the 3 benzil dioximes only I forms the complex $\text{N}_1(\text{C}_6\text{H}_5\text{O}_2\text{N}_2)_2$ by substitution of 1 oxime H atom. The reduction of III by nascent H or PCl_5 to IV shows the presence of 2 extra nuclear O atoms, confirms the formula

given and excludes the formula $\text{O O N CPh CPh N CPh CPh}$. The reactions de-

scribed in the present paper as well as those already known of glyoximes show that it is impossible to generalize, e. g., α -p-tolil dioxime and anisil dioxime which have many properties in common with I, yield on oxidation with $\text{K}_2\text{Fe}(\text{CN})_6$ no trace of III.

C. C. DAVIS

Aminomalonate ester. Action of alkyl iodides and bromides on diethylsodium-aminomalonate. R. LOCQUIN AND V. CERCHER *Bull. soc. chim.* [4], 47, 1377-80 (1930).—This is an introductory paper stating that $\text{NH}_2\text{CNa}(\text{CO}_2\text{Et})_2$ reacts with alkyl iodides and bromides to give $\text{RC}(\text{NH}_2)(\text{CO}_2\text{Et})_2$ in 50% yield. These substances are basic; they form cryst. compds. with KOCN and PhNCO . Previous work on $\text{NH}_2\text{CH}(\text{CO}_2\text{Et})_2$ is reviewed.

I. M. LEVINE

Some alkylaminomalonate esters and their immediate derivatives. V. CERCHER *Bull. soc. chim.* [4], 47, 1381-5 (1930).—By the action of 1 mol. of iso- $\text{C}_4\text{H}_9\text{I}$ on 1 mol. of $\text{NH}_2\text{CNa}(\text{CO}_2\text{Et})_2$ either in alc. or in iso- $\text{C}_4\text{H}_9\text{ONa}$ there was obtained 50% of *di-Et* α -aminoisobutylmalonate (I), $\text{C}_8\text{H}_{17}\text{C}(\text{CO}_2\text{Et})_2\text{NH}_2$, b₁₀ 136° , n_D^{20} 1.4371, d_4^{25} 1.0109, M. R. 59.93 (calcd. 59.75). Treated with KCNO in AcOH I gave the corresponding urea, $\text{C}_8\text{H}_{17}\text{C}(\text{CO}_2\text{Et})_2\text{NHCONH}_2$, m 194° , and with PhCNO in Et_2O the phenylurea, $\text{C}_8\text{H}_{17}\text{C}(\text{CO}_2\text{Et})_2\text{NHCONHPh}$, m 128° , in 70% yield. On mixing 2 g. of I with 8 g. of water satd. with NH_3 at 0° , together with a few drops of EtOH and shaking vigorously, the diamide, $\text{C}_8\text{H}_{17}\text{C}(\text{CONH}_2)_2\text{NH}_2$, m 153° , was formed on 24 hrs. standing. *Di-Et* aminoallylmalonate (II), $\text{CH}_2=\text{CHCH}_2\text{C}(\text{CO}_2\text{Et})_2\text{NH}_2$, was prepd. from $\text{CH}_2=\text{CHCH}_2\text{Br}$ and $\text{NH}_2\text{CNa}(\text{CO}_2\text{Et})_2$, II, b₁₀ $134-5^\circ$, n_D^{20} 1.4442, d_4^{25} 1.039, M. R. 55.00 (calcd. 54.67); urea, m $173-4^\circ$; phenylurea, m 113° . Attempts to make *di-Et* aminobenzylmalonate, $\text{PhCH}_2\text{C}(\text{CO}_2\text{Et})_2\text{NH}_2$, resulted in noncryst. oils, decomposed on heating, which gave with KCNO the urea, m 234° , and with NH_4OH the diamide, m 156° .

C. R. ADDINALL

Preparation of α -amino acids by hydrolysis of alkylaminomalonate esters. R. LOCQUIN AND V. CERCHER *Bull. soc. chim.* [4], 47, 1386-9 (1930).—Alkylaminomalonate esters are hydrolyzed either by means of superheated steam at 150° in an autoclave or by refluxing with 10% HCl , to yield α -amino acids, CO_2 and ROH . The use of steam is preferable unless the resulting acid is insol. in H_2O . $\text{NH}_2\text{CH}_2\text{CH}(\text{CO}_2\text{H})$ (I) (9.5 g. from 25 g. of $\text{NH}_2\text{CH}(\text{CO}_2\text{Et})_2$), leucine, $\text{PhCH}_2\text{CH}(\text{NH}_2)\text{CO}_2\text{H}$ and $\text{CH}_3\text{CH}(\text{CH}_2\text{CH}(\text{NH}_2)\text{CO}_2\text{H})$ were obtained in this way. In the prepn. of I, the compd. I.HCl , m 185° , was isolated.

I. M. LEVINE

The structure of some organic molecules. O. HASSEL AND E. NAESHAGEN *Tids. Kemi. Bergesen* 10, 126-7 (1930), cf. C. A. 24, 2346.—A preliminary notice concerning work on dipole moments, and discussing in this connection *meso*- and *di* forms. Racemic and active dichlorosuccinic acids gave almost identical values for the dipole moment (2.93×10^{-18}), the *meso*-form gave lower values (2.47×10^{-18}). β -Hexachlorocyclohexane (*trans*) has a center of symmetry, and dipole moment 0, while α -hexachlorocyclohexane (*cis*) has dipole moment 2.20×10^{-18} esu.

G. R. YONE

Preparation of bromoacetyl sugars and of acetoglucals. P. A. LEVENE AND ALBERT L. RAYMOND *J. Biol. Chem.* 90, 217-50 (1931).—The prepn. of BrCH_2CO derivs. of sugars can be greatly simplified by dissolving the reaction product resulting from the

action of HBr and Ac_2O on the sugar or of HBr in AcOH on the pentaacetate in C_2H_5 concd the soln at 40–50° under reduced pressure and renewing the C_2H_5 until all the excess reagents are removed. The $\text{BrC}_2\text{H}_5\text{CO}$ deriv generally crystallizes directly in the distg flask. The procedure in the prepn of acetoglucals is also much simplified by the use of C_2H_5 ; it forms an azeotropic mixt. with the AcOH used as a solvent when treating the bromoacetyl sugar with Zn dust. A. P. LOTTROP

Crystalline α - and β -methyl *d*-gulonides. HORACE S. ISBELL. *Proc Nat Acad Sci* 16, 699–703 (1930). cf. *C. A.* 25, 1223 —Although a sirup of Me gulonides prep'd by condensation of the pure sugar with MeOH in the presence of 1% HCl could not be crystd directly, the isomeric gulonides were sep'd. by means of their cryst. addn compds with CaCl_2 . By removal of CaCl_2 with $(\text{CO}_2\text{Ag})_2$ the pure cryst. Me gulonides were prep'd and their parent CaCl_2 compds repred. as a check on their original purity. The new compds described are α -Me *d* gulonide $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$, $[\alpha]_D^{25}$ 67°, $[\text{M}]$ 22,900, (α -Me *d*-gulonide) $\cdot \text{CaCl}_2 \cdot 3\text{H}_2\text{O}$, $[\alpha]_D^{25}$ 83°, $[\text{M}]$ 23,000, α -Me *d*-gulonide H_2O , decramps 77°, $[\alpha]_D^{25}$ 109°, $[\text{M}]$ 23,100, β -Me *d* gulonide $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$, ca. 155°, moderately sol in abs. alc., $[\alpha]_D^{25}$ –45°, $[\text{M}]$ –15,640, (β -Me *d*-gulonide) $\cdot \text{CaCl}_2$ very slightly sol in abs. alc., $[\alpha]_D^{25}$ –65°, $[\text{M}]$ –16,200, β -Me *d* gulonide, m 176° (uncor.) $[\alpha]_D^{25}$ –83°, $[\text{M}]$ –16,100. The mol. rotations of all the α forms are in good agreement with one another as are those of the three β -forms. The difference between the mol. rotations of the α - and the β -gulonides is 39,200 which compares favorably with the normal value 37,000. P. C. HOCKETT

β -Lactose. JOHN GORTOV. *Svensk Farm Tids* 34, 665–7 (1930) —Lactose (100 g) dissolved in 60 cc. H_2O over a free flame and set aside at 105° to crystallize gave 90% of β -lactose with $[\alpha]_D^{25}$ 39°, changing to 53.6° in 3 hrs. The $[\alpha]_D^{25}$ of the original lactose was 83.7°. The β -soln. when made alk. with NH_3 gave $[\alpha]_D^{25}$ 55.2° and the α 52.7°. A. R. ROSE

Recent syntheses in the sugar group. AMÉ PICTET. *Atti III congresso naz. chim. pura applicata* 1930, 39–41 —A brief review is given of the efforts to synthesize sugar with special emphasis on the methods involving condensation of the anhydrosugars and on the synthesis of sucrose. By heating starch with glycerol, 2 series of anhydrides (hexa-, tri- and diglucosans) are obtained which may be considered as polymers of glucosan. The 1st series of compds. is unaffected by I and the last member is an anhydride of maltose. The members of the other series are colored by I and the simplest is an anhydride of isomaltose. The members of the 1st series polymerize readily under the influence of light, giving substances colored by I. The triglucosan of the 2nd series in concd. aq. soln. protected from light, in time deposits a colloidal substance colored violet by I which is strikingly analogous to starch itself. Its acetate shows by f p. methods that it must be a dodecaglucosan, $(\text{C}_6\text{H}_7\text{O}_5)_n$. R. C. H.

Inulin. X. H. PRINGSHEIM, J. REILLY, W. G. HENSEL, W. BURMEISTER, P. P. DONOVAN and MISS N. HAYES. *Ber* 63B, 2636–42 (1930); cf. *C. A.* 24, 3995 —Berner (*C. A.* 24, 5730) has criticized the observation that inulin dissolves in liquid NH_3 , fused AcNH_3 , and (as its acetate) in AcOH with a mol. wt. corresponding to a difructose anhydride, and that by heat disaggregation and boiling of its acetate with PhSO_3H , as well as by soln. in AcNH_3 , it is converted into an inulan of the same mol. wt. B. ascribes the depressions observed to impurities such as H_2O or alc. in the inulin and AcNH_3 in the inulan. The authors have shown by the C_4H_8 reaction that the inulin they used was alc. free and that the 1% of H_2O in the air-dried product is generally given off in 0.5 hr. at 78° in *vacuo* over P_2O_5 . The inulan on distn. from alk. soln. gave no more color with Nessler reagent than did dried H_2O . The mol. wt. detns. were repeated independently at Berlin and Cork with due regard to these sources of error and the results confirmed by measurements in purified HCONH_2 . Observing the precautions specified by Freudenberg (*C. A.* 24, 1346) values corresponding to $2 \times (\text{C}_6\text{H}_7\text{O}_5)_2$ were in most cases obtained for inulin in both HCONH_2 and AcNH_3 , although double this value was obtained in 2 cases, whether this was due to previous treatment in the prepn. and purification or to the fact that natural inulin, as believed by Schlubach, is not a homogeneous substance remains to be tested. The inulan in H_2O soln., when sufficiently d.d. again gave values corresponding to the $2 \times \text{C}_6$ -stage. Drying has some influence on its original H_2O -solv. but even after especially vigorous drying at 110° it showed, after soln. in hot H_2O , a mol. wt. (387) not very much higher than 2×162 . Schlubach and Elsner and Jackson and Goergen have observed that a small amt. of glucose is formed along with the fructose in the acid hydrolysis of inulin. The present authors also observed, by differential titration according to Bertrand on the one hand and Willstätter-Schudel on the other, an aldose content higher than could be accounted

for by exptl error (Ch I, 100 70 94) Expts on pure fructose under the conditions of the Willstätter titration showed that it is attacked to the extent of about 2.8%, the aldose content in the acid hydrolyzate of inulin thus being reduced to about 5%. Taking this titration error into account, no glucose could be found in the inulinase hydrolyzate, thus confirming Bourquelot and Bridel. This, however, does not prove that inulin contains no glucose, for the mass of fructose obtained by the ferment hydrolysis, as detd with Lehling soln, was only 60% even after weeks and months.

C A R

Methylated tri- and tetra-saccharides from cellulose and starch. KARL FRIDRICH BERG AND KARL FRIDRICH *Naturwissenschaften* 18, 1114(1929).—Recently cryst decamethyl β methylcellobioside, b.p. 216–20° was prepd by acetylation and methylation of cellulose (C A 25, 25). The cellobiose deriv, b 60° lower is also obtained by the same distn. The methyltetrose b.p. 255–55° and is also obtained in crystals with all the properties of a tridecamethyl β methylcellobioside. Both compds crystallize easily in water or petr ether even when the MeO content is 1.2% too low. Repeated methylation of the cryst preps gives an analytically correct and uniform product with the expected mol wt in camphor. The methylated trisaccharide, m 118°, has α_D^{25} –17° in water the methylated tetrasaccharide, m 139°, α_D^{25} 14.6°. From the products of acetylated starch corresponding, but noncrystallizable, fractions were obtained after methylation a methylated trisaccharide, α_D^{25} 133° (water), mol wt. 613 (calcd 678), and a methylated tetrasaccharide, α_D^{25} 149° (water), mol wt 812 (calcd 862). These products of starch are the α glucosidic analogs of the methylated oligosaccharides from cellulose, the former having MeO in α as well as β positions. From demolition kinetics it was deduced (cf preceding abstr) that starch as well as cellulose has a chain structure. The present data are addl evidence. The same structure is required for inulin contrary to the opinions of Pringsheim (following abstr), Schlubach (C. A. 25, 215) and Staudinger (C A 25, 290). B J C v d H.

Chemistry of starch. XXIV. New polysmyloses. I. HANS PRINGHEIM, ALFRED WIENER AND ALEXANDER WEIDINGER. *Ber* 63B, 2628–36(1930), cf C A 24, 3996.—When AcNH_2 , which, it had been found, can be used as a f p solvent for complex saccharides such as inulin and glycogen, was tried with α tetraamylose (I), it gave values for the mol wt corresponding, not to a tetra- or diamylose, but to a hexose anhydride (162). In view of the possible effect of the high temp of molten AcNH_2 (above 80°) on the labile polyamylose, the mol-wt detns were repeated in HCONH_2 but with the same results. At first higher values (never below and often above 2 \times 162) were obtained, but this was traced to impurities in the HCONH_2 (Schering Kahlbaum, m –5° to –7°, with about 1% HCO_2H). When this had been purified until it was acid free and m. 1.8–2°, it always gave for 1 mol wt corresponding to a glucose anhydride. Pptn of the HCONH_2 soln with alc. gave 80% of a cryst substance (II) sepg from hot aq alc with 1 mol H_2O in column-shaped prisms having, when dried, the compn and mol wt. (in H_2O) of a glucose anhydride and giving with I-KI the long green needles with metallic luster characteristic of α amyloses. II is therefore designated as α -amylosan. The mother liquor from the II, freed of alc. and HCO_2NH_2 in *vacuo* and pptd with $\text{EtOH-Et}_2\text{O}$, gave an amorphous product (III), sepg from dil alc in rhombic tables with 1 mol H_2O , which could best be differentiated from II by means of I-KI which gave brown red prismatic columns like β polyamyloses. III, which is called β -amylosan, also has the compn $\text{C}_6\text{H}_{10}\text{O}_5$, but gives a mol wt of 2 \times 162 in both H_2O and HCONH_2 . β -Heptaamylose (IV) showed a mol wt of 2 \times 162 in both AcNH_2 and HCONH_2 and when pptd from HCONH_2 with $\text{EtOH-Et}_2\text{O}$ yielded 90% of III. Like the polyamyloses II and III form with org solvents difficultly H_2O sol addn products by means of which they can readily be isolated from aq solns. The II and III were further characterized by means of their acetates. In AcOH the acetate of III was dispersed nearly to the C_6 stage in 24 hrs. On heating, the rotations of aq solns of II and III change to a final const. value (II, from 150° to 128°; III, from 179° to 150°) and recrystn of the resulting products from dil alc. gives well-crystd substances, α - (V) and β -isoamylosan (VI). Neither V and VI nor their I and II addn products differ visibly in cryst form from II and III and their corresponding derivs. In HCONH_2 V and VI give the expected mol. wts and after pptn their rotations are found unchanged. In H_2O , however, they show a pronounced tendency to aggregate, gradually passing after several days at room temp into a colloid-dispersed state. II, heated 6 hrs on the H_2O bath in HCONH_2 , changed to a substance (VII) sepg from aq alc. in 6-sided tables and yielding an especially characteristic I addn product in black green prisms. I, II, III and IV in aq CaH_2N (7.3) all showed

the same specific rotation (137°) II, $[\alpha]_D^{20}$ $141.8-9^\circ$ (H_2O) III, $[\alpha]_D^{20}$ $173-81^\circ$ (H_2O) α -Triacetilamylon, $[\alpha]_D^{20}$ $120-4^\circ$ ($CHCl_3$), mol wt. 256-307 (AcOH) β -Triacetilamylon, $[\alpha]_D^{20}$ $129-32.5^\circ$ ($CHCl_3$), mol wt. 322 (AcOH) V, $[\alpha]_D^{20}$ $127-33^\circ$ (H_2O) VI, $[\alpha]_D^{20}$ $150-2^\circ$ (H_2O), 185.5° (H_2O -EtOH, 1:1) C. A. R.

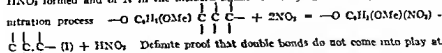
Lignin and cellulose. XIV. The hydrolysis of polysaccharides. KARL FREUDENBERG, WERNER KUTN, WALTER DIER, FRITZ BOLL and GUSTAV STEINBRUNN. *Ber. 63B*, 1510-30 (1930), cf. *C. A.* 24, 1846.—In the series of researches relating to the kinetics of the degradation of polysaccharides, the measurements on disaccharides and other sugar derivs formed a communication preliminary to those on cellulose and starch (*C. A.* 23, 176). Starch is now investigated under the conditions used with cellobiose and maltose. For cellulose 5% H_2SO_4 was used, soln. and hydrolysis followed just as with starch and cello-dextrin (the biosan of Hess). The detn. of the aldehyde groups freed was made with 0.1 N I. The Cu method does not work because cellobiose reduces more Cu than 0.5 glucose, the 2nd glucose group is attacked after destruction of the 1st. The velocity const. for the hydrolysis of cellobiose by 50% H_2SO_4 at 18° was found to be 1.07×10^{-4} (μ in min.) by means of well-agreeing titrations and polarimetric observations. With cellulose the degree of hydrolysis as detd. polarimetrically is in advance of the degree as detd. by titration with I. The object was to find which of the possible courses of reaction (see last reference hypotheses II, III and V) best suits the data as obtained by the latter method. Of the consts. assumed, k_1 is known from the velocity of hydrolysis of cellobiose. k_2 is the initial velocity const. of the reaction in all 3 cases. It is found either by extrapolation to an infinitely short time from the mean velocity const. P_{∞} or is detd. from k_1 and any one value of P (P_{∞} was chosen, i. e., P for $\alpha = 0.50$) with the aid of the tables calcd. from one of the 3 fundamental formulas expressing the relation between k_1 , k_2 and P_{∞} . The degree of agreement between the entire courses of the empirical and theoretical curves shows the applicability of the fundamental relation on which the formula used is based. Hypothesis II (k_1 only for cellobiose different from k_2) is eliminated while hypothesis III (k_1 for disaccharide and trisaccharide) gives good and hypothesis V (biosan conception) fairly good agreement. Formula III is the better expression, especially for the final stage of the reaction. The fission of cellulose at 30° gives entirely analogous results. Hydrolysis of biosan acetate (Hess and Friese, *C. A.* 21, 174). Biosan acetate was crysd. from $CHCl_3$ -MeOH, extd. with the latter and the Ac groups removed according to Fischer and Zemplén, the H_2O -sol. portions of the free dextrin were used. $[\alpha]_D^{20}$ in 2 N NaOH -9.95° . Fission at 18° and 30° . k_1 is as large as in the case of cellulose and here also hypothesis III is better suited than V to explain the kinetics. Hydrolysis of starch— k_1 (maltose) at 18° 1.60×10^{-4} . Kahlbaum's sol. starch, $[\alpha]_D^{20}$ in formamide (0.5%) 195° . k_1 by extrapolation $= 1.07 \times 10^{-4}$, by hypothesis III, found 1.19×10^{-4} . The curve constructed according to this agrees well with that found empirically. Hypothesis II is eliminated here but not the biosan assumption (V). Similar facts hold good for the hydrolysis at 30° . Hydrolysis consts. of disaccharides.—Solns. were 0.1 N in 2 N H_2SO_4 at 70° (μ in min.): gentobiose 3.06, levoglucosan 6.0 trimethyl levoglucosan 1.0, Me mannoside ($[\alpha]_D^{20}$ 79° in H_2O) 1.8×10^{-4} in good agreement with similar measurements of Moolwyn Hughes (*C. A.* 22, 1888, 23, 5099, 5393, 24, 2154). The consts. given by von Meyer, Hopff and Mark (*C. A.* 23, 4451) for the in version of maltose are not in agreement. Optical Rotations.—In connection with the polarimetric detn. of the degree of degradation, it is to be borne in mind that the rotatory power unlike the I value does not lie around the half point between glucose and cellulose (see above). An effort is made to reach a better approximation by assuming the rotations of the tri- and tetrasaccharides to correspond to their position between the disaccharide and polysaccharide. The values calcd. from assumption III are in better accord with the actual detns. than those from V but a close agreement between the values so calcd. and the α -values found titrimetrically is not obtainable. The former III remain somewhat behind the detd. values which indicates that the rotations of the tri- and tetrasaccharides are somewhat higher than those indicated by the new assumption. Thus the deviations may be explained without difficulty. In contrast the biosan theory requires exact applicability of the derived formulas. Large and const. deviations indicate more than mere exptl. errors. From the dependence of the velocity consts. of hydrolysis upon temp. the heats of activation are derived for the fission of the compds. and the thermal proportionality factors. The heats of activation of starch and maltose are not significantly different from one another while those of cellulose and cellobiose are well distinguished, the former being smaller. That, nevertheless, the velocity const. for cellulose is smaller is to be ascribed to the much

smaller thermal proportionality factor. Similar, but less pronounced, are the relations with starch and maltose. The kinetic considerations exclude a 2nd type of linking along with the cellobiose linking. Since fully methylated cellobiose decomposes only into 2,3,6-trimethylglucose, the only alternative lies between the α - and β -linkings on C atom 4. A disaccharide more difficultly hydrolyzed than cellobiose or one equally easily split, along with cellobiose, would be comprehensible but one more easily hydrolyzed like maltose would make itself noticeable through a more rapid course of reaction. With cellobiose the extrapolated initial I consumption indicates a chain length of 30 to 40 glucose residues (mol wt around 5000 to 7000). The authors explain the deviation from the lower value of Bergmann and Machemer (*C. A.* 24, 2877) by the fact that in the free dextrin used, the lower water sol parts were removed and these must have been present in the brown acetate used by Bergmann. With cellulose, a parallel situation is found, giving a chain length of 50 glucoses. Cellobiose is close to cellulose also as indicated by its lower rotation. The chain of cellulose is assumed to be actually longer; not small admixtures of lower carbohydrates can raise the I value. The very different properties of cellobiose from those of cellulose (sol., viscosity) are to be attributed to the proportionate length of chain: the former having been partially shortened, the entirely degraded fragments especially being absent. From the I requirement, starch indicates a chain length of 25-30 glucose residues. The sol starch used was perhaps partly degraded already. A strict literal application of assumption III forces us to make a fundamental distinction between the di- and trisaccharides and the tetrasaccharides. If the middle linking within a long chain always reacts according to k_1 , the outer linking according to k_2 , then formula III would be a 1st but only a poor approximation. Necessary modifications would have to be made if higher fission consists of intermediate values between k_1 and k_2 were to be ascribed to the outer linkings. It would be conceivable that the glucoses should be split off more rapidly at 1 or both ends (perhaps according to k_1) while the rest hydrolyzed more or less according to k_1 . The detn. thus far of the max. yield of cellobiose depended upon the assumption of similar stability of the linkings throughout the chain. If this proves incorrect, then the cellobiose resulting during all stages of degradation would have to remain under 67%, which agrees with a series of new results (50-60%). With starch, it is more difficult to decide regarding the small difference between the initial and final velocities, but the complete analogy between the degradation and the optical changes makes a chain structure seem likely here also. The high yields of acetobromomaltose which Karrer (*C. A.* 16, 63) obtained by the AcBr degradation of starch are explained by the fact that that substance is attacked more slowly by AcBr than the chain compds. The diastatic degradation is explained by a similar linking of all the glucose residues of the starch chain, but with a series of accessory assumptions confirmed by kinetic observations. A chain formula is also required for maln. The newly detd. facts are reconcilable with the chain formulas and a uniform interpretation is possible only with this conception.

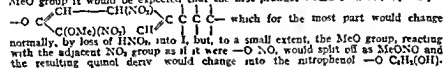
R. C. HOCKETT

Lignin and cellulose. XV. Lignin and nitrogen peroxide. KARL FRETZENBERG AND WALTER DURR. *Ber.* 63B, 2713-20, (1930), cf preceding abstr., Schaarschmidt, *Z. anorg. Chem.* 42, 618 (1929).—After it had been found that the reaction of Br on lignin consists in substitution and not addn., it remained to be detd. whether lignin behaves toward nitrating agents also as an aromatic substance. The action of HNO_3 on lignin is not clear cut, the reaction does not stop at a definite stage but comes to a standstill only when products of deep-seated degradation, which are no longer characteristic, have been formed. Better results were to be expected of NO_2 , which with aromatic compds. gives substitution derivs., preceded by addn. of 2 mols NO_2 at the double bond, with subsequent elimination of HNO_2 . In preliminary expts. in a N-filled app. in ice, connected with a manometer to record changes in pressure, the pressure, within a few min. after bringing the NO_2 (mixture of NO_2 and N_2O_4) and lignin together, fell far lower than in blank expts. with NO_2 alone and remained const. for a while, but after about 0.5 hr. it rose slowly for 3 days, owing to the evolution of NO. Attempts to follow the reaction quantitatively failed because of the extraordinary adsorbing power of the lignin and the complex equil. between NO , NO_2 , N_2O_4 , HNO_2 , HNO_3 and H_2O . There is no doubt that the gradual increase in pressure is due to a secondary reaction, evidently an oxidation process, and that only the first rapid phase (decrease in pressure) marks the course of the reaction sought. This is even more evident when methyl-lignin is used instead of lignin. Etherification seems to block the point of attack for the oxidation reaction; with methyl-lignin the pressure falls to a min. within a few min. and then remains unchanged for days. The expts. were then conducted under conditions which permitted of converting the nitrous gases in the reaction vessel after

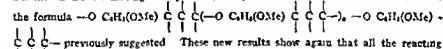
the nitration into a mixt. of nitrite and nitrate which was then analyzed by the method of Meisenheimer and Hein. The gas remaining after absorption of the nitrous gases in aq. NaHCO_3 was pure NO , no org. oxidation products (especially no $(\text{CO}_2\text{H})_2$) were found in the NaHCO_3 soln and no CO_2 was formed in the nitration. The only cleavage product found was MeOH . The amounts of NO_2 used, of NO , HNO_3 and HNO_2 formed and of N in the nitrated lignin or methylgignin correspond best to a



all or only to a small extent is afforded by the fact that the nitrolignin and nitromethylgignin can still be brominated (nitrolignin takes up about 0.7 atom Br per atom of N) and that bromomethylgignin can be nitrated. The cleavage of MeOH (20-33% of that originally present in the lignin) had already been observed in the bromination, as it occurs to the same extent whether lignin or methylgignin is used, there is no doubt that it is the original MeO group in lignin that is split off. This cleavage of MeOH is explained on the basis of Wieland's theory that the substitution of NO_2 is preceded by an addn. Assuming, arbitrarily, that the NO_2 group enters the p -position to the MeO group it would be expected that the first product would be a di- NO_2 compd



$(\text{NO}_2) \text{ C} \text{ C} \text{ C}-$. Dimethyldehydrovanillin, $[\text{5,2,3 OHC}(\text{MeO})_2\text{C}_6\text{H}_3]$, in solid amorphous form, treated with NO_2 like lignin, is similarly nitrated, with elimination of 2.5 mols MeOH . Veratric acid, both in solid form and in PhNO_2 is nitrated but splits off only traces of MeOH . Veratrylcellulose (18.6% MeO) takes up 12.5% NO_2 but almost without loss of MeOH . F and D feel that the results of bromination and nitration show that lignin is predominantly an aromatic substance and confirm the



groups are free since the lignin reacts with the completeness of a permutoid. A further example (the mercuriation of lignin) is to be described shortly. Lignin differs from a directed permutoid like siloxene in that the structural elements apparently extend in 3 dimensions entirely without any order, resulting in an extraordinary development of surface, as indicated by the strong adsorbing power. A further peculiarity is that the structural elements are apparently not all entirely identical, although derived from a basic form closely related to coniferyl alc or guaiacylglycerol. Presumably there are differences in the aliphatic chains, and furthermore a small part of the aromatic nuclei are present in the form of the piperonyl residue. The size of the particles formed by condensation also seems to vary widely in one and the same prepn. As a result, the reactions of lignin are by far not nearly so homogeneous as those of permutoids of regular structure. C. A. R.

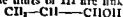
Constitution of cellulose xanthate. III. TH. LIESER. *Ann* 483, 132-43 (1930), cf. C. A. 23, 3672.—Treatment of methylcellulose with a mixt. of AcOH and Ac_2O contg. H_2SO_4 at 25-30° for 7-20 days gives about 50% of cellobiose octaacetate, indicating that some of the glucose anhydride units in the original product are not methylated. No methylcellobiose acetate could be detected, it is probably decomposed during acetolysis. C. J. WEST.

Action of mercury salts on acetohalogenosugars. IV. Direct preparation of alkyl biosides of the α -series. GÉZA ZEMPLEN and ARPAD GERECES. *Ber* 63B, 2720-9 (1930), cf. C. A. 24, 5730.—It had been found that heptaacetyl α -alkylcellobiosides can be obtained in faultlessly pure condition by the action of $\text{Hg}(\text{OAc})_2$ on acetobromocellobiose (I) in the presence of the appropriate alc., but the results were not always reproducible, especially with the Et compd., and the quality of the product varied

widely with the amt of the alc used. Systematic expts were therefore carried out in which only the quantity of alc used was varied. The mixt of 1 (10 g), $\text{Hg}(\text{OAc})_2$ (2 g) and 1 tOH was boiled 2 hrs in 10 cc C_2H_5 , the cooled reaction mixt washed 4 times with H_2O , the C_2H_5 soln dried with CaCl_2 , evapor in vacuo, then twice with EtOH and the residue dissolved in 50 cc hot alc and allowed to stand 6 hrs at room temp. The product, recrystd once more from alc, was then tested for reducing power, rotation in CHCl_3 and m p. The results showed that whether the product solvated shall be the α or the β form can be controlled by the proper choice of the quantity of alc used. The α Et deriv can be obtained with certainty with an excess in the neighborhood of 100% of EtOH, even with a 200% excess the product is still a quite pure α deriv but the yield is no longer satisfactory (with a 100% excess the purity and the yield are as good as in the isomerization of the β compd with TiCl_4 and the procedure is materially simpler). When the excess of alc exceeds 300%, however, the reducing power of the product again increases and between a 300 and 400% excess of alc there is a sharp transition in favor of the β form. With 10 g 1, a difference of 0.74 g in the quantity of EtOH used sufficed to change the rotation of the product from 50.86° to -18.20° . With 150 EtOH the α deriv can be obtained more easily than with 1 EtOH, a 40–200% excess of the 150 EtOH giving an optically very pure product in good yields, the transition in favor of the β form with a further increase in the quantity of alc, however, is even more marked than with EtOH. With Pr, Bu, iso Bu, sec Bu, sec Am, hexyl and PhCH₂CH₂ alcs, a 100% excess of the alc often gave the α deriv of highest rotation. To obtain the β forms, a large excess of the alc must be used. With the lower alcs, it is advantageous to use the alc itself as the reaction medium. The β derivs can often also be obtained readily with $\text{Hg}(\text{CN})_2$, instead of $\text{Hg}(\text{OAc})_2$, but attempts to prep the α compds in this way met with but little success. The highest $[\alpha]_D$ in CHCl_3 observed for the α and β forms, resp., of the heptaerythritolketohexonides were: Me —, -25.0° , Et 37.2° , —, Pr 58.70° , —, iso-Pr 59.29° , —, 22.7° , Bu (m 172°) 52.40° , -21.60° , iso Bu (m 171°) 45.51° , —, 23.04° , sec-Bu (m 193°) 55.76° , -23.20° , sec Am (m 193°) 52.21° , —, hexyl (m 182°) 53.42° , -21.37° , β -phenylethyl (m 207°) 51.16° , -25.25° . C. A. R.

New substances formed by molds. NADINE WIJCKMAN, *Ann* 485, 61–73 (1931). — The action of *Penicillium glaucum* on sucrose in new glass vessels gives a mixt of gluconic acids 1 (I) and 2 (II), sep'd by crystn from EtOH. I, $\text{C}_6\text{H}_{12}\text{O}_7$, m 202° , is optically inactive, contains no MeO group, gives no color with FeCl_3 and from its behavior on titration appears to contain 1 or more lactone groups. I is not esterified with MeOH-HCl , with CH_3N_3 , N is evolved, giving a product boiling at 210° in high vacuum, it could not be crystd. Ac deriv, m 175° , H_2 deriv, m 170° , NO_2 deriv, m 159° . Reduction of I with Zn and AcOH gives the compd $\text{C}_6\text{H}_{12}\text{O}_6$, m 217° . Distn of I in N at 215° gives the unsat'd ketone (III), $\text{C}_6\text{H}_{10}\text{O}_6$, b.p. 120° , which absorbs 2 Br and 2 I (dinitrophenylhydrazones, red, m 173° , semicarbazone, m 183°) and the compd, $\text{C}_7\text{H}_{12}\text{O}_6$, m 171° . Oxidation of III gives the compd $\text{C}_6\text{H}_{10}\text{O}_7$, analyzed as the *alg salt*, II, $\text{C}_6\text{H}_{10}\text{O}_8$, m 180° , dist'd undecompd in a high vacuum and titrates as a 4 basic acid. I is probably a HO deriv of II. C. J. WEST.

Xylan. J. REILLY, P. P. DONOVAN AND MISS K. BURNS, *Proc Roy Irish Acad.* 39B, 505–14 (1930), cf. C. A. 23, 337; 24, 831, 3990. — The work on the effect of soln of poly saccharides in AcNH_3 (I) and HCONH_2 (II) was extended to the pentosans. Abs 1 EtOH pptd from a 0.5% soln of xylan (III) in I a substance which was sol in cold H_2O while moist but which on drying became horny and insol in cold H_2O but sol on short heating to 50° . This ppt (called xylanin) (IV) possessed the same rotatory power as the original III, $([\alpha]_D^{20})_{\text{conc}} -111.1^\circ$ to -113° , and analysis showed its empirical formula to be $(\text{C}_5\text{H}_8\text{O}_4)_n$. F-p detns of the mol wt on III in I at 0.451% concn gave the values 146 and 147; at 0.955%, 195 and at 1.1%, 231. These values point definitely to depolymerization with formation of the mols $\text{C}_5\text{H}_8\text{O}_4$ and $(\text{C}_5\text{H}_8\text{O}_4)_2$. Mol-wt detns in H_2O on dried IV from I gave the values 310, 348 and 485 at 0.73% concn, and 353, 412 and 480 at 0.93% concn. Similar detns on IV from II gave 341, 383 and 453 at 0.925% concn, and 369, 415 and 513 at 1.049% concn. The values indicate the formula $(\text{C}_5\text{H}_8\text{O}_4)_n$ for IV with increasing assoc on standing. I and II may act as imino acids to produce depolymerization or it may be that the fundamental units of III are assoc through their residual valences and that these valences are altered when the III is dissolved in I or II. If the units of III are linked through auxiliary



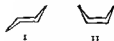
valences, the structure of the unit should be:



C. H. PEET

Resolution of some ester acids of the γ -truxillic acids into the optical components FRIEDRICH SCHENCK. *Ber.* 63B, 2706-12(1930)—It has thus far not been possible to resolve the mono esters of α truxillic and β truxinic acids into their optical components Stoermer and Fretwurst showed, however, that the γ -truxillic acids are readily resolved with morphine. The present work was undertaken to det whether the monoesters of γ -truxillic acid are also resolvable. To obtain the active monoesters, the following active amidic acids, some of which had already been described by Fretwurst were prepd: d -Me (+)- γ truxillic amide (I), l Me (-)- d -amide (II), l Et (-)- d -amide (III), l Pr (-)- d -amide (IV), d Bu (+)- l -amide (V). These were converted with N_2O_5 in AcOH into the following truxillic monoesters: Me (+)- d γ -truxillate (VI), Me (-)- l truxillate (VII), Et (-)- l truxillate (VIII), Pr (-)- l truxillate (IX), Bu (+)- d truxillate (X). As standards of reference were taken Fretwurst's active amidic acids. The l form corresponds to the negative, the d form to the positive rotation. When such an acid (e g, the d amidic acid) is esterified it becomes, as regards the ester group, the l form, i e, the d amidic l -ester, for if the amide group should be saponid and the remaining ester group replaced by NH_2 , the l amidic acid should result. Attempts to effect such a replacement have hitherto failed. As with the amidic acids, the d forms of the monoesters have a pos, the l -forms a neg rotation. The direction of the rotation of VII and VIII, which was not detd directly, was deduced from the m ps with the active compds obtained by resolution (see below). The inversion of the rotation on esterification of the amidic acids noted by Fretwurst was again observed, as was also the increase in the magnitude of the rotation. To obtain the dl monoesters, γ truxillic anhydride was boiled with an excess of alc, and about 0.5 mol Na_2CO_3 , and the resolution was effected by bringing together hot satd alc. solns. of 1 mol each of the monoester and morphine, the salts of the (-)- l Me, (-)- l Et, (-)- l Bu and (+)- d Pr esters crystg out while the antipodes remained in soln and sepd. only on cautious evapn. With brucine, the antipodes formed the difficultly sol salts (the Bu ester could not be resolved with this alkaloid). The magnitude of the rotation increases with increasing mol wt. of the alc., and the m ps of the dl esters decrease while those of the active esters increase with increasing mol wt., with the exception of the Me ester dl -isomers: Me, m 183.5-4°, Et, m 173-4.5°, Pr, m 170-1°, does not depress the m p of the Et ester, Bu, m 142.3°, seps from H_2O as a hydrate, m 62-3° II, m 153-4° III, m 153-4° IV, m 150-60°, $[\alpha]_D^{20}$ -17.7° (all rotations measured in Me_2CO) V, m 141.5-3°, $[\alpha]_D^{20}$ 18.33°. d Bu γ truxillate, from the acid with $BuOH \cdot H_2SO_4$, m 85-6° VII, from II with N_2O_5 or from the dl ester with morphine, m 145.5-7°, a mixt with VI m 178-83° VI, from I and N_2O_5 or from the dl -ester with brucine, m. 145.5-7°, $[\alpha]_D^{20}$ 6.48° (prepd by 1st method), $[\alpha]_D^{20}$ 4.37° (prepd by 2nd method) VIII, m 142.5-3°, $[\alpha]_D^{20}$ -16.48°, m 173° when mixed with the (+)- d -ester (from the dl -ester with morphine), m 142.5-3°, $[\alpha]_D^{20}$ 19.14° IX, m 148-9°, $[\alpha]_D^{20}$ -21.35° when prepd from IV with N_2O_5 , $[\alpha]_D^{20}$ -22.45° when prepd from the dl -ester with brucine, m. 168-70° with the (+)- d -ester (from the dl -ester with morphine), m 148-9°, Bu (-)- l γ truxillate from the dl -ester with morphine, m 151-2°, $[\alpha]_D^{20}$ -25.70° X, m 151.2° $[\alpha]_D^{20}$ 26.14° C. A. R.

The structure of the cyclohexane molecule O HASSEL AND H KRINGSTAD *Tidskr. Aem. Bergsten* 10, 128-30(1930)—Since satd hydrocarbons have practically 0 dipole moment, a distinction between the possibilities I and II for cyclohexane by dipole measurement is not possible. X ray investigation, with a Debye camera.



CaH_2 frozen with solid CO_2 -acetone mixt., and Fe radiation gave the following data. Lattice cubic $a \approx 8.41 \approx 0.02 \text{ A U}$, 4 mols. per unit cell, d 0.933. The space group is probably T_d^2 or O_h^4 . These results indicate the form of the cyclohexane mol. to be I. G. R. Y.

Two syntheses of heptamethylenimine ADOLF MÜLLER AND PAUL BLEIER. *Monatsh.* 56, 391-406 (1930)—Reduction of suberone isoxime (a ketoheptamethylenimine) with Na and $EtOH$ gives 1.5% slightly impure heptamethylenimine (I), b 162-4°, d_4^{20} 0.869 n_D^{20} 1.5262 (chloroaurate suffers 173°, m. 188°, picrate, m 147-8°, N -Bz deriv. b 196-7°). Oxidation of the N -Bz deriv. with $KMnO_4$, at 65° gives 41.3% of β benzamidoheptanoic acid dl m. mpous, m 86° and 90°, this also results by benzoylating the amino acid obtained by heating suberone isoxime with 20% HCl at 150°. Hydrolysis of the reaction product of α - γ -dibromoheptane and β - $MeC_6H_4SO_3NH_2$ with $EtOH$ KOH by HCl at 16° yields 0.8% of slightly impure I and α , γ -diamino-

heptane. Improved methods are given for the prepn of suberic acid, suberone and its isoxime C J WEST

Migration of amino groups from one to three in arylamine derivatives of diarylarylethynylcarbinols. Constitution of resulting products. JOSEPH RONIN *Compt rend* 191, 790-2(1930), cf C A 24, 78—PhNH₂ with diphenylphenylethynylcarbinyl chloride (Ph₂CCIC CPh) (I) gives a yellow product, Ph₂C CHCPh NPh (II), m 199-200°. A colorless intermediate product, C₁₂H₉N (III), m 87-8°, has been isolated. III isomerizes easily into II by the action of heat in the presence of PhNH₂, HCl, but not in the presence of PhNH₂. II has not been converted into III. III is more basic than II, forming an isolable HCl salt. With MeOH in dil H₂SO₄ II gives β-phenylbenzylacetophenone, while III gives diphenylethynylcarbinyl Me ether. III is probably Ph C(NHPh)C CPh. These relationships hold in a variety of similar cases where an aryl amine with a diarylarylethynylcarbinyl chloride gives 2 isomers, the unstable product (a) of the type of formula II being white, and the stable product (b) of the type of formula III being yellow. I with *o*-toluidine gives (a), m 139-40°, and (b), m 102-3°, with *m*-toluidine (a), m 117-8° (b), m 139-40°, with *p*-toluidine (a), m 115-6° (b), m 174-5°, with *o*-anisidine (a), m 140-1° (b) not crystd., with *p*-anisidine (a) not crystd., (b) m 149-9°. Diphenyl *p*-tolylethynylcarbinyl chloride with PhNH₂ gave (a), m 118-9°, (b), m 162-3°. Diphenyl 3-naphthylethynylcarbinyl chloride with PhNH₂ gave (a), m 140-7°, (b), m 149-50°. Diphenyl *p*-bromophenylethynylcarbinyl chloride with PhNH₂ gave (a), m 151-2°, (b) m 138-9°. T H RIDER

The action of hexahalogenated benzenes on Grignards. J F DUBAND AND LAI WAL HSUEY *Compt rend* 191, 1460-3(1930)—C₆Cl₆ reacts with neither MeMgI nor PhMgBr. C₆Br₆ or C₆I₆ with MeMgI gives C₆Me₆, m 157-8° and with PhMgBr C₆Ph₆, m 266°, b 418-20°. V F HARRINGTON

Chlorination of *o*- and *p*-nitrotoluene. 3,4,6-Trichloro-2-nitrotoluene and 2,6-dichloro-4-nitrotoluene. LEOPOLD F LEVY AND HENRY STEPHEN *J Chem Soc* 1931, 76-9—*o*-O₂NC₆H₄Me (27.4 g) and 5 g SbCl₅ acid with Cl₂ give 5 g of the 3,4,6-Cl₃ deriv, pale yellow, m 93-4°, further nitration gives the 2,5-di-NO₂ deriv, m 227.5° (cor.). Reduction with SnCl₂ and HCl gives 3,4,6-trichloro-*o*-toluidine, m 89°, Ac deriv, m 199°; Bz deriv, m 230° *p*-O₂NC₆H₄Me, chlorinated as above, gives the 2,6-Cl₂ deriv, m 63-4°; HNO₃ and H₂SO₄ give the 3,4-di-NO₂ deriv, m 130°. 2,6-Dichloro-*p*-toluidine, m 55°; Ac deriv, m 215°; Bz deriv, m 179°. Both of the new Cl derivs resist oxidation after heating with KMnO₄ for 6 days on the water bath.

C J WEST

A new color reaction of aromatic amines. I DE PAOLINI *Gazz chim ital* 60, 859-62(1930)—Br₂O₂ (I) is an extremely sensitive reagent for detecting aromatic amines, for it will give a brown red or violet color reaction with the smallest trace of the amine. This color results from the oxidation of the amine to a quinone, and the reaction has the great advantage over previous color reactions (cf. *Ann* 27, 376(1853), *Compt rend* 111, 975(1890), *Ber* 27, 3264(1894)) that the oxidation takes place in neutral alc. soln. The reaction of I with aliphatic amines has been studied by Gambharjan (cf. C A 20, 372), who found the reaction to be: I + R₁R₂NH → Br₂OII + R₁R₂NOBz. P. has, however, found that with primary amines another aminolysis takes place, thus: I + RNH₂ → BrO₂H (II) + RNHBr. II oxidizes in turn a 2nd mol of RNH₂. E. g., PhCH₂NH₂ (III) and I yield PhCH₂NHBr (IV), Br₂II (from the subsequent oxidation of III by II) and BrONH₂CH₂Ph (from combination of III and BrO₂H, the latter the reduction product of II). Practically, I is added to alc. III until no more is dissolved, and the soln is dild with water, which causes pptn first of BrOH and then of IV. I reacts with other RNH₂ compds, e. g., with alc. NH₂OH or H₂NNH₂. H₂O it forms BrNHOH and [BrNH]₂, resp. Gas is evolved in these reactions, because II (which is formed in each case) oxidizes the excess NH₂OH and H₂NNH₂. H₂O to the extent that the mols are destroyed. In the presence of water and of oxidizable substances, II is formed by the action not only of primary amines, but also of secondary and tertiary amines, as a result of a hydrolysis similar to that with alk. hydroxides, i. e., I + NaOH → BrONa + II, as was shown in the present work by using dil alc. solns contg indigo. The great ease of oxidation of aromatic amines to quinones by I probably results from the intermediate formation of II, which in the nascent state is extremely active; in fact the reaction is only feeble in anhyd solvents, e. g., Et₂O or AcMe. The color reaction was obtained with PhNH₂, *o*-, *m*- and *p*-MeC₆H₄NH₂, asym *o*-xylydine, *p*-xylydine, *p*-anisidine, *o*- and *p*-H₂NC₆H₄OH, *o*-, *m*- and *p*-C₆H₄(NH₂)₂, benzidine, PhNHCH₂Ph, PhNHEt, PhNHMe, PhNMe₂, PhNEt₂, *o*- and *p*-MeC₆H₄NMe₂, PhNEtCH₂Ph, MeNPh, *p*-HOC₆H₄NH₂, *o*- and *p*-C₆H₄NH₂, β-C₆H₄NHPh, Ph₂C(NMe₂)₂, *o*-, *m*- and *p*-H₂NC₆H₄CO₂H, 4,3-Me-

(H₂N)C₆H₄CO₂H, 3-bromo-6-aminocuminic acid, 3-aminocuminic acid, 2,4,5-BrMe-(H₂N)C₆H₄CO₂H, *o*-, *m*- and *p*-ClC₆H₄NH₂, 2,4-Cl₂C₆H₃NH₂, *o*-, *m*- and *p*-BrC₆H₄NH₂, and *p*-IC₆H₄NH₂. In all cases it is best to add the amine to a suspension of I in LiOH, under which conditions a brown red or violet color appears even when cold, but more rapidly when hot. The reaction is being studied further. C. C. DAVIS

Asymmetric nitrogen atom. LVIII. Decomposition of active and inactive quaternary ammonium nitrates under the influence of amines; also a contribution to the knowledge of solvate formation by ammonium salts. E. WERKIND AND F. FEISTEL. *Ber* 63B, 2743-53 (1930), cf. C. A. 23, 5171.—It had been shown that the decomposition of MePh(PbCH₃)₂(PhMeNCH₂CH₂ONO₂) (I) in CHCl₃ is greatly accelerated by amines, secondary bases having a greater influence than primary bases and tertiary bases having little or no effect, whether the I is dissolved directly in the amines or treated in CHCl₃ with small quantities of the bases. The reaction was explained as consisting in a primary formation of a solvate (indicated by an increase of the sp. rotation, among other things) which, because of its relative instability, brings about a decomposition of the nitrate into a ditertiary base and PbCH₃NO₂ and the latter finally combines with the excess of amine to form a new salt in which an amine H atom is replaced by PbCH₃. The renewed salt formation as the result of solvolysis is indicated by a renewed increase in elec. cond. which during the preceding phases decreases along with the rotation. As these expts. were made with a salt (I) which was not completely stable, especially in CHCl₃, they have been repeated with simple active NH₄ nitrates (especially MePh(PbCH₃)₂(CH₃CHCH₃)NNO₂ (II)) and a perchlorate, MePh(PbCH₃)₂ClO₄ (III) (the nitrate is not crystall., as also with a no. of other amines. To obtain an idea of the qualitative tendency of the halides and nitrates of these NH₄ compounds they were tested for a change in cond. with time. The results in CHCl₃ and on addn. of PhNH₂, resp., were as follows (+ indicates decompn., - no decompn., n.t. not tested) MePh(PbCH₃)₂(CH₃CHCH₃)NI +, n.t., nitrate (II) -, + MePh(PbCH₃)₂(iso-Bu)NI +, n.t. nitrate -, + MeEtPh(PbCH₃)NI +, n.t., nitrate -, + MePh(PbCH₃)₂(PhCH₃)NI +, n.t. nitrate -, + PhCH₃C(CH₃)MeEt(CH₃CHCH₃)₂Br₂ (IV) -, -, dimnate -, - MeEtPhNI (V) -, -, perchlorate (III) -, -. The halides which do not decomp. in CHCl₃ (IV, V) also do not react on addn. of PhNH₂, and the same is true for the corresponding nitrate and perchlorate. Only those halides decomp. which contain certain groups (Ph and PhCH₃ or CH₃CHCH₃). Since the nitrates which decomp. on addn. of PhNH₂ react essentially alike, kinetic measurements were made first on MePh(CH₃CHCH₃)NNO₂ (VI) and various amines by the method used before (measurement of the decrease in rotation and cond. with time). With 0.1 g. VI and 50 mole amine in 10 cc. CHCl₃, the following half time values (min.) and av. velocity consts. resp. were obtained: Primary amines: 140-Am 23, 0.002; 140-Bu 45, 0.0072; Pr 47, 0.0085; Bu 88, 0.0064; PhCH₃ 90, 0.0037; Ph 250, 0.0032; Secondary amines: Et 24, 0.0039; Et(PhCH₃) 42, 0.0176; (PhCH₃)₂ 80, 0.0177; MePh 90, 0.0077; EtPh 150, 0.0055; Ph(PhCH₃) 195, 0.0014; Ph₃ —. The tertiary PhNMe₂ produced no measurable change in 18 hrs. The relationships between basicity and reaction velocity thus brought out will be discussed elsewhere in connection with further exptl. data. As regards the mechanism of the reaction and especially the differing behavior of primary and secondary amines on the 1 hand and tertiary amines on the other, it is to be remembered that, according to the hypothesis suggested in connection with the observations on I, the presence of at least one H on the N of the added amine is necessary for the reaction to take place, for it is assumed that a H atom combines with the NO₂ ion to form HNO₂ and is replaced by one of the α -carbon residues, leaving a tertiary base. NR'R''R'''NO₂ + HNR' = NR'R''R''' + R'NHR' + HNO₂. The HNO₂ may be bound to 1 or both of the bases. It was found that when VI is dissolved in fused *p*-toluidine the cooled solidified mass yields MeC₆H₄NH₂·HNO₂ decompn. 197°. The work on I had also led to the conclusion that solvates are formed primarily in the action of amines on the nitrates. The simple method then used for showing their formation (placing the nitrate alone, the amine alone and a mixt. of nitrate and amine in weighed boats and heating them in a high vacuum at gradually increasing temps. up to 111°) was employed in the present work. With those nitrates which decomp. on addn. of the amine the pure nitrate, rema. unchanged while the mixt. of nitrate and amine lost wt. until the residue weighed less than the original nitrate, although the presence of amine (PhNH₂) could still be detected by the isonitrate reaction. With III and the nitrate corresponding to IV, the wt. of the residue never fell as low as that of the original nitrate or III in the mixt., there is nothing to prove, however, that these compds. do not also form solvates which are relatively stable as compared with those of the other

nitrate. With tertiary bases (PhNMe_3 , PhNMe_3I) there is no further loss in wt. after the residue has reached the wt. of the original nitrate. No solvate formation between VI and CHCl_3 could be detected by this method. The $[\alpha]_D^{20}$ of VI in various solvents (Me_2CO 56°, abs. alc. 67°, PhNHMe 73°, PhNH_2 75°, CHCl_3 80°) is smallest in those solvents which are indifferent as regards their influence on the decompn. (Me_2CO , alc.) and highest in those which are known to effect the decompn. In CHCl_3 , which according to the heating method forms no solvate, the rotation, although higher than in PhNHMe and PhNH_2 , is independent of the concn., whereas in PhNHMe and PhNH_2 it increases with increasing concn. That no solvation occurs in CHCl_3 in the absence of an added base was confirmed by l_p measurements on II in CHBr_3 , the mol wts. found (674 v, 583, at concns. of 2.75, 1.83) indicated the usual assocn. of NH_2 salts in CHBr_3 but gave no evidence whatever of simultaneous solvation. In p -toluidine, which causes the decompn. of II in CHCl_3 , the mol wt. of II decreases with the length of time it is kept in the molten toluidine until a const. value of 73.40 (calcd. 300.7) is reached, while III, which is not decomposed by the toluidine in CHCl_3 , gives mol wts. greater than the calcd. but decreasing with increasing concn., indicating the presence of solvates. C. A. R.

Diphenylamine derivatives. K. DZIEWONSKI AND M. RUSSACKI *Bull. intern. acad. Polonaise* 1929A, 50, 27.—The conclusions of Merz and Weith (*Ber.* 5, 283) concerning the proportion of mono- and di- SO_3H acids formed by the action of concd. H_2SO_4 on Ph_2NH are confirmed. With ClSO_3H in PhNO_2 below 0°, however, the initial product is the unstable, cryst. $\text{ClSO}_3\text{H} \cdot \text{NHPh}_2$, which at higher temps. readily decomposes, giving nuclear substitution products in accordance with the scheme $\text{Ph}_2\text{NH} \cdot \text{SO}_3\text{Cl} \rightarrow \text{Ph}_2\text{NSO}_3\text{H} \rightarrow \text{PhIN}(\text{C}_6\text{H}_5)_2\text{SO}_3\text{H}$, the proportion of mono- and di- SO_3H acids produced depending on the mol. proportion of ClSO_3H used. Thus with 0.5, 1.0, 1.5 and 2.0 mols. of ClSO_3H for each mol. of Ph_2NH the proportions of mono- and di- SO_3H acids and of unchanged Ph_2NH are, resp., 23, 5, 70, 33, 22, 40, 34, 50, 16, and 0, 91-100, 0 mols. % Hence with 0.5 mol. of the reagent, diphenylamine-4-sulfonic acid (aniline salt, m. 230.5°) is readily sepd. from the reaction mixt. as its Mg salt. Nitration of 4- $\text{PhNH}(\text{C}_6\text{H}_5)_2\text{SO}_3\text{Na}$ with 1 mol. of HNO_3 (d. 1.48) in AcOH gives the 4-nitro-deriv. (*Mg salt*) converted by hydrolysis with concd. HCl in a sealed tube at 190° into 4- $\text{PhNH}(\text{C}_6\text{H}_5)_2\text{NO}_2$ (Goldberg, Ger. pat. 185,663(1907)), and further nitrated to Na 2,4-dinitrodiphenylamine-4-sulfonate. Excess of HNO_3 converts 4- $\text{PhNH}(\text{C}_6\text{H}_5)_2\text{SO}_3\text{Na}$ into the 2,4,6-trinitro deriv. identical with a specimen obtained by the condensation of 2,4,6- $\text{Cl}_3\text{C}_6\text{H}_2\text{NO}_2$ and sulfanilic acid (Turpin, *J. Chem. Soc.* 59, 717(1891)) and hydrolyzed by concd. HCl at 180-200° to 2,4,6-(O_2N) $_3\text{C}_6\text{H}_2\text{NHPh}$. The action of Br (2 mols.) on a suspension of 4- $\text{PhNH}(\text{C}_6\text{H}_5)_2\text{SO}_3\text{Na}$ in CCl_4 at 45-50° converts it into its 2-Br deriv., while with 6 mols. of Br at 60-70° the product is Na 2,4,2'-tribromodiphenylamine-4-sulfonate (*Mg salt*), converted by excess of HBr at 60° into (2,4- $\text{Br}_2\text{C}_6\text{H}_3$) $_2\text{NH}$ and by dil. HCl at 190° into 2,4,2'-tribromodiphenylamine, m. 94°, while the action of HNO_3 (d. 1.48) in AcOH at 165-90° yields 2,4,2'-tribromo-6'-nitrodiphenylamine, m. 215°, together with the corresponding 4',6'-di-nitro deriv., m. 179°, and Na 2,4,2'-tribromo-6'-nitrodiphenylamine-4-sulfonate. Diphenylamine-4,4'-disulfonic acid (*Mg salt* + 2.5 H_2O , and aniline salt, m. 234°) is best obtained by the action of an excess of ClSO_3H on Ph_2NH in PhNO_2 at 110-15°. By reactions similar to those described in the case of the mono- SO_3H acid, its Na salt is converted into the 2-nitro- and 2,2'-dinitrodiphenylamine-disulfonic acids (converted into (2- $\text{O}_2\text{N}(\text{C}_6\text{H}_4)_2\text{NH}$, 2,4,2'-trinitrodiphenylaminesulfonic acid), and [2,4-(O_2N) $_2\text{C}_6\text{H}_3$] $_2\text{NH}$, small quantities of the 2,4,6,2',4'-pentanitro- and 2,4,6,2',4',6'-hexanitro derivs. being formed as by-products of the nitrations). Similarly, bromination of 4,4'- $\text{NH}(\text{C}_6\text{H}_4\text{SO}_3)_2\text{Ba}$ converts it into the 2,2'-di-Br deriv., + 0.5 H_2O , together with some 2-Br derivs., and, ultimately, into (2,4- $\text{Br}_2\text{C}_6\text{H}_3$) $_2\text{NH}$. B. C. A.

An improved method of preparation of substituted amides and hydrazides. SOHRAB M. MISTRY AND PRAPILLU C. GUHA. *J. Indian Chem. Soc.* 7, 783-7(1930).—The condensation between divers types of amines, diamines, hydrazines, thioethercarbazides, etc., on the one hand and amides such as $\text{CO}(\text{NH}_2)_2$, $\text{CS}(\text{NH}_2)_2$, PhCONH_2 , $\text{C}_6\text{H}_5(\text{CO})\text{NH}$ on the other, can be brought about, with excellent yields, by refluxing a mixt. of 1 mol. of the amide and 2 mols. of the amine with sufficient AmOH until the evolution of NH_3 is negligible. With NH_2NH_2 , H_2O and $(\text{NH}_2)_2\text{CO}$ (I), $\text{NH}_2\text{NHCNH}_2$ is obtained to the exclusion of the expected $(\text{NH}_2\text{NH}_2)_2\text{CO}$. From I and the requisite amine the following known compds. were prepd.: $(\text{PhNH})_2\text{CO}$, $(m\text{-MeC}_6\text{H}_4\text{NH})_2\text{CO}$, $(p\text{-MeC}_6\text{H}_4\text{NH})_2\text{CO}$, $(m\text{-NO}_2\text{C}_6\text{H}_4\text{NH})_2\text{CO}$, $(p\text{-NO}_2\text{C}_6\text{H}_4\text{NH})_2\text{CO}$, $(\beta\text{-C}_6\text{H}_4\text{NH})_2\text{CO}$, $(\alpha\text{-C}_6\text{H}_4\text{NH})_2\text{CO}$, $(p\text{-NHAcC}_6\text{H}_4\text{NH})_2\text{CO}$, $(\text{PhCH}_2\text{NH})_2\text{CO}$, *o*- and *p*-phenyleneurea

$\text{NH C}_6\text{H}_5$, C_6H_5 , NH CO (from benzidine) $(\text{PhNHNH})_2\text{CO}$ With I , m $\text{NH}_2\text{C}_6\text{H}_4\text{OH}$ gave $(m\text{-HOC}_6\text{H}_4\text{NH})_2\text{CO}$, m 215° , and $p\text{-NH}_2\text{C}_6\text{H}_4\text{OH}$ similarly yielded d - p -hydroxyphenyl urea, m 288° . The condensation with I converted $p\text{-NH}_2\text{C}_6\text{H}_4\text{COOH}$ into 2,4-dihydroxy

tetrahydro-1,2,3,4-quinazoline, $\text{C}_6\text{H}_5\text{CH(OH)NHCH(OH)NH}$, m 353° (decompn) D - p -nitrophenylcarbohydrazide, $(p\text{-NO}_2\text{C}_6\text{H}_4\text{NHNH})_2\text{CO}$, m 381° and d - p -bromophenyl carbohydrazide were made from $p\text{-NO}_2\text{C}_6\text{H}_4\text{NHNH}$, and $p\text{-BrC}_6\text{H}_4\text{NHNH}$, resp. The treatment of $\text{NH}_2\text{NHCSNHPh}$ with I by this method gave 3 compds: 1-phenyl 2-phenyl

amino-5-thiol-1,3,4-triazole, $\text{PhNC(NHPh)NNC}_6\text{H}_4\text{SH}$, m 210° , 2,5-endoxy-1,3,4-triazole,

$\text{NH C}\cdot\text{N}\cdot\text{N}\cdot\text{C}\cdot\text{O}$, m 250° , and 2 keto-5-phenylamino-2,3-dihydro-1,3,4-thiodi-

azole, S CO NH N CNHPh , m 246° . $p\text{-AcNH C}_6\text{H}_4\text{NH}_2$ (II) formed $\text{AcNH C}_6\text{H}_4\text{NHCONHPh}$ with NH_2CONHPh and $(\text{AcNH C}_6\text{H}_4\text{NH})_2\text{CS}$ with $(\text{NH}_2)_2\text{CS}$ (III). III gave $(\text{PhNH})_2\text{CS}$ with PhNH (IV) but formed $\text{NH}_2\text{NHCSNHPh}$ with PhNHNH (V). $(\text{NH}_2)_2\text{CO}$, condensed with II, IV and V to yield $(p\text{-AcNH C}_6\text{H}_4\text{NHCO})_2$, m above 370° , $(\text{PhNHCO})_2$ and $(\text{PhNHNHCO})_2$. II, IV and V gave $(p\text{-AcNH C}_6\text{H}_4\text{NHCO})_2\text{CH}_2$, m 235° ; $\text{C}_6\text{H}_5(\text{CONHPh})_2$ and $\text{C}_6\text{H}_5(\text{CONHPh})_2$ with Cl^+ , $(\text{CONH})_2$, and when condensed with $(\text{CH}_3\text{CONH})_2$, yielded $(p\text{-AcNH C}_6\text{H}_4\text{NHCOCH}_3)_2$, m 347° (decompn), $(\text{PhNHCOCH}_3)_2$, and $(\text{PhNHNHCOCH}_3)_2$. AcNHPh was obtained in 80% yield from AcNH_2 and IV, on complete removal of AmOH and crystn of the residue from hot H_2O . C R ADDINALL

Chlorination of iodophenols. III, Chlorination of o -iodophenol. SAMUEL BUCHAN AND HAMILTON MCCOMBIE *J Chem Soc* 1931, 137-44, cf *C A* 7, 1708--a IC. H_2O (I) in CCl_4 gives on chlorination a sequence of unstable iododichlorides and decompn products very similar to those obtained from the p isomer. I acetate, oil, dichloride, m $92\text{--}3^\circ$ (decompn), unstable I benzyl ether, oil, the dichloride is unstable I benzoate, m 34° , dichloride, m $98\text{--}101^\circ$ (decompn), stable I phenyl carbamate m $121\text{--}2^\circ$. Cl in CHCl_3 gives the dichloride, m 125° (decompn) of 2-iodo-phenyl 2,4-dichlorophenylcarbamate, m 145° , its constitution was proved by its synthesis from the 2 components. A 10% soln of I in CCl_4 satd with dry Cl at 0° , gives 4-chloro-2-iodophenol (II), m 78° , its constitution was established by synthesis from 4,2- $\text{Cl(O}_2\text{N)}\text{C}_6\text{H}_4\text{OH}$ whose benzyl deriv, m $84\text{--}5^\circ$. II acetate oil, the dichloride decompn rapidly. II benzyl ether, m 60° , dichloride, decompn $95\text{--}7^\circ$, decompn slowly over 4-5 days. II benzoate, m 88° , dichloride, stable m $103\text{--}4^\circ$ (decompn). II phenylcarbamate, m 128° , Cl in CHCl_3 gives the compd $\text{C}_6\text{H}_5\text{O}_2\text{NCl}_2$, m 146° . Further chlorination of II in CCl_4 gives 4,6-dichloro-2-iodophenol (III) m 62° , this is more easily obtained by rapid chlorination of I. III acetate, m 59° , the dichloride decompn in 5-6 days. III benzyl ether, m 62° , the dichloride decompn over several days. III benzoate, m 83° , chlorination does not give the dichloride. III phenyl carbamate, m 181° , no dichloride pptd with Cl but HCl was evolved. Further chlorination of III gives a small amt of the tri Cl deriv, m $52\text{--}4^\circ$. 2 Iododiphenyl ether, m $55\text{--}6^\circ$, results from the NO_2 deriv by reduction and the Sandmeyer reaction, the dichloride, m $81\text{--}2^\circ$ (decompn), is unstable and decompn in 2 days. o -2' Iodophenoxyacetophenone m 123° from I and $\text{BrC}_6\text{H}_4\text{Cl}$ dichloride, m 98° (decompn), moderately stable. I and COCl_2 in pyridine give 4-chloro-2-iodophenyl carbamate, m 88° , dichloride m 108° (decompn), stable. I p -toluenesulfonate, m 80° , dichloride, m $95\text{--}7^\circ$ (decompn), stable. I allyl ether, oil, Cl gives the compd $\text{C}_6\text{H}_5\text{OCl}_2$, m 84° (decompn) slowly decompn in the air. I dibromopropyl ether, oil, the dichloride, m 118° (decompn) decompn slowly in 4 weeks, giving 4-chloro-2-iodophenyl dibromopropyl ether m 52° . I Et carbonate, oil, a dichloride could not be prepd. Chlorination of I in CCl_4 at $60\text{--}70^\circ$ gives chloranil and hexachlorocyclohexadienone. p - $\text{C}_6\text{H}_4\text{OH}$ and ICl give 4-chloro-2,6-diiodophenol, m $100\text{--}7^\circ$, the dichloride is unstable. Chlorination of a 10% soln of PhNHCO_2Ph in CHCl_3 at 0° gives 4-chlorophenyl 2,4-dichlorophenylcarbamate, m 157° , its structure was established by condensing $p\text{-ClC}_6\text{H}_4\text{OH}$ with 2,4-dichlorophenylcarbamide, m 61° . C J WEST

Chlorination and nitration in methanol as a medium. E PLAZEK ROZCNIK Chem 10, 761-76 (778 German) (1930).—In order to investigate the possibility of introducing MeOH as a solvent and diluent in the chlorination and bromination of various compds, PhOH , m -cresol, PhOMe , 2- and 3- $\text{HOC}_6\text{H}_4\text{CO}_2\text{H}$, benzoic acid, 4- $\text{NO}_2\text{C}_6\text{H}_4\text{OH}$, 1,2-acetylphenylenediamine, 4-sulfophenol, 4- $\text{NO}_2\text{C}_6\text{H}_4\text{NH}_2$, 2- $\text{NO}_2\text{C}_6\text{H}_4\text{NH}_2$,

2 and 4- $\text{H}_2\text{NC}_6\text{H}_4\text{CO}_2\text{H}$ were chlorinated and PhOH, *m*-cresol, 4- $\text{ClC}_6\text{H}_4\text{OH}$, 2- $\text{ClC}_6\text{H}_4\text{OH}$, veratrole and $p\text{-C}_6\text{H}_4(\text{OMe})_2$ nitrated in MeOH as a medium. MeOH was found to be a good solvent in chlorinations, equal in value to AcOH; it possesses the advantage of greater volatility and can be more easily regenerated. In some cases the influence of temp and diln on the course of chlorination was investigated and found to be insignificant. Analogous expts with EtOH showed that this solvent is much inferior to MeOH in chlorination and quite unsuitable in nitration. Nitration in MeOH could be performed with success only in the case of PhOH and its derivs, especially 2 and 4- $\text{ClC}_6\text{H}_4\text{OH}$, veratrole and $p\text{-C}_6\text{H}_4(\text{OMe})_2$. J. KUDRKA

New method of reducing aromatic nitro compounds. THEODORE R. DE KIEWITZ AND HENRY STEPHEN. *J. Chem. Soc.* 1931, 82-4.—Cryst. SnCl_4 is dehydrated with Ac_2O and the nitro compd added and heated until soln results, in this way, e. g., PhNO_2 gives nearly quant. $p\text{-ClC}_6\text{H}_4\text{NHAc}$. Difficulty was experienced in the isolation of the anilides resulting from the reduction of *m* substituted NO_2 compds and it was found necessary to isolate the free bases by steam distn in alk. soln. A 30% yield of the anilides may be obtained if the residue after removal of the AcOH is extrd. with hot C_6H_6 . 3-Chloro-4-acetamidophenol, m. 186° , and 3-chloro-4-acetamidophenyl acetate, m. 127° , result by the reduction of 3,6 and 3,4- $\text{Cl}(\text{O}_2\text{N})\text{C}_6\text{H}_3\text{OH}$. In *p*-substituted NO_2 compds no chlorination takes place except in the case of $p\text{-ClC}_6\text{H}_4\text{NO}_2$ which, only if HCl is present during the reduction, gives 2,4- $\text{Cl}_2\text{C}_6\text{H}_3\text{NHAc}$. The reduction of $o\text{-O}_2\text{NC}_6\text{H}_4\text{OH}$ and certain of its derivs yielded in each case a *N*-Ac deriv., although sufficient Ac_2O was present to acetylate the HO group. α -Methylbenzoxanines may be intermediate products in their reduction. The reduction products isolated from a no. of compds are listed. C. J. WEST

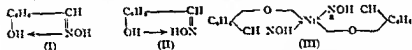
The production of *p*-aminophenol. JULIUS ALTMAYER. *Metallurgie* 20, 2803-4 (1930); 21, 27-8 (1931).—A review is given of the various methods devised for the production of $p\text{-NH}_2\text{C}_6\text{H}_4\text{OH}$ (I), and of the reactions in which I is a product. Short descriptions are given of the reduction of $p\text{-NO}_2\text{C}_6\text{H}_4\text{OH}$ by means of metal and acid combinations, by P, I and HCl, by the use of Na_2S , and by catalytic hydrogenation, and of the employment of metal acid combinations and of Na_2S in the reduction of $p\text{-NOC}_6\text{H}_4\text{OH}$. Prepn. depending on the cleavage of azo compds include the reduction of $p\text{-dihydroxyazobenzene}$ by SnCl_2 in acid soln, Zn and NaOH and alk. Na_2S . Benzotriazophenol (II) in Na_2CO_3 , treated with H_2S , gives a mixt. of PhNH_2 and I. Other methods for the cleavage of II include treatment with alc. NH_3 , extrd. with H_2S , reduction with metal-acid combinations and fusion with NH_4NHPh . By the reduction of 4,4'- $\text{HOC}_6\text{H}_4\text{N}=\text{NC}_6\text{H}_4\text{SO}_3\text{Na}$ with Fe and H_2SO_4 , 80% of I and 95% of $\text{NH}_2\text{C}_6\text{H}_4\text{SO}_3\text{H}$ are obtained. The reduction can also be carried out with Fe and NaCl. The diazo dye formed by coupling 2 mols. of PhOH with tetrazodiphenyl can be cleaved by catalytic hydrogenation into $(\text{C}_6\text{H}_5\text{NH}_2)_2$ and I. Under the heading of various methods of prepn. and formation of I are listed the change of $p\text{-ClC}_6\text{H}_4\text{OH}$ into I by NH_4OH and CuSO_4 ; the conversion of PhN_3 by the action of H_2SO_4 in the presence of AcOH; the oxidation of PhNH_2 by $\text{K}_2\text{Cr}_2\text{O}_7$ and of $\text{HO}_3\text{SC}_6\text{H}_4\text{NH}_2$ by MnO_2 and H_2SO_4 ; the transformation of PhNH_2OH on heating with acid, and a similar transformation following the soln. of PhNO_2 in a soln. of S in 35-40% oleum at 20° ; the reaction of NOCl with C_6H_6 in the presence of AlCl_3 ; the action of H_2O_2 on PhNHMeBr in Et_2O ; the hydrolyses of *N*-phenyl α -indole oxime and phenolphthalein oxime by 25% H_2SO_4 ; the cleavage of nitrones and the distn. of 5,2- $\text{H}_2\text{N}(\text{HO})\text{C}_6\text{H}_3\text{CO}_2\text{H}$. The original method of Gattermann for the prepn. of I by the electrolysis of a soln. of NO_2Ph in H_2SO_4 has been modified in many ways. The various changes involving cell construction, electrodes, solns, temp., agitation, c. d., time, by-products of the reaction, etc., are summarized. C. R. A.

Reaction of *p*-anisidine and ethyl methylacetacetate. WM. O. KERMACK AND JAMES F. SMITH. *J. Chem. Soc.* 1931, 221-2.— $p\text{-MeOC}_6\text{H}_4\text{NH}_2$ (I) and $\text{AcCHMeCO}_2\text{Et}$ (II) appear to give a compd., $\text{C}_{11}\text{H}_{13}\text{O}_3\text{N}_2$, formed from 2 mols. I and 1 mol. II, m. 51° ; this is confirmed by analysis, its gradual decompn. in a stoppered bottle, its behavior with Ac_2O and mol.-wt. detn. C. J. WEST

Preparation of an α -aminophenolsulfonic acid. JAKOB POLLAK AND KARL DEUTSCHER. *Monatsh.* 56, 365-80 (1930).— $(2\text{-H}_2\text{NC}_6\text{H}_4)_2\text{S}_2$ is decomd. by ClSO_3H at -20° and by fuming H_2SO_4 under various conditions. Phenylenediazosulfide is unaffected by fuming H_2SO_4 (25-70% SO_3) at $160-70^\circ$. $o\text{-H}_2\text{NC}_6\text{H}_4\text{SH}$ could not be sulfonated. 4,3- $\text{Cl}(\text{O}_2\text{N})\text{C}_6\text{H}_3\text{SO}_3\text{K}$ and Na_2S_2 in 95% EtOH give *K* 2,2'-dinitrodiphenyl disulfide-4,4'-disulfonate; in H_2O reduction occurs with either Na_2S or Na_2S_2 . The Ba salt of 4,3- $\text{Cl}(\text{O}_2\text{N})\text{C}_6\text{H}_3\text{SO}_3\text{H}$ and *K* xanthate in H_2O give EtOH, Et xanthate, CS_2 , and the Ba salt (I) of 2,2'-dinitrodiphenyl sulfide-4,4'-disulfonic acid (II), crystg. with 5 mols. H_2O , with AcONH_4 , only a small amount of the sulfide is produced. II

and PCl_5 give the *dichloride* of II, orange-yellow, m 195° , crystg with 0.5 mol C_6H_5 ; oxidation of the chloride with fuming HNO_3 gives 2,2'-diaminodiphenyl sulfone-4,4'-disulfonic acid, whose Pb salt crystallizes with 3 mols H_2O and whose Ba salt also crystallizes with 3 mols H_2O , this also results by the oxidation of I. Reduction of I with Sn and HCl gives 2,2'-diaminodiphenyl sulfide-4,4'-disulfonic acid, crystallizes with 3 mols H_2O and readily oxidizes in the air. The diazo compd gives a red diazo dye with $\beta\text{-C}_{10}\text{H}_7\text{OH}$. Treatment of the reaction product from 4,3 $\text{Cl}(\text{O}_2\text{N})\text{C}_6\text{H}_4\text{SO}_3\text{Na}$ and $\text{Na}_2\text{S}_2\text{O}_3$ with PCl_5 also gives the chloride; when the K sulfonate is treated similarly Na 2,2'-diaminodiphenyl sulfide-4,4'-disulfonate is obtained 4,3 $\text{Cl}(\text{O}_2\text{N})\text{C}_6\text{H}_4\text{SO}_3\text{K}$ and KHS , heated in H_2O , the soln evapd to dryness and the residue reduced with Sn and HCl , give 75% of 2-aminothiophenyl-4-sulfonic acid. NaNO_2 gives a diazo compd which does not couple with $\beta\text{-C}_{10}\text{H}_7\text{OH}$. With the K sulfonate there results K 5-nitrophenothiazine 3,8-disulfonate, crystg with 2 mols H_2O . Numerous attempts to prep the acid with other S -contg salts were unsuccessful. C. J. WEST

Coordination compounds of oximes. II. Nickel and cobalt compounds of *o*-hydroxybenzaloxime. OSCAR L. BAARV *J Chem Soc* 1931, 105-7; cf *C A* 24, 5296—*o*- $\text{HOC}_6\text{H}_4\text{CH NOH}$ (I) has all the properties of a coordination compd and probably should be formulated as I, rather than II, because it yields a NOAc deriv



with Ac_2O at room temp. The Ni salt, dark green needles, and the Co salt, dark brown powder, are of the type III. No evidence of the β isomers could be obtained. Under the conditions specified by Ephraim (*C A* 24, 5065) for the detn of Cu by I, Ni salts give some ppt., the presence of salts of weak acids in any quantity would also invalidate the method. C. J. WEST

Action of mustard oils on oximes. AN ORREGIA AND C. V. GEORGHIOU *J prakt Chem* 128, 239-319 (1930)—*o*- PhCH NOH (I) and PhNCS (II) at room temp in a sealed tube give COS , PhCN and $(\text{PhNH})_2\text{CS}$, in boiling C_6H_6 the reaction products are the same while if the 2 compds are heated without a solvent to 80° COS is evolved and S and $(\text{PhNH})_2\text{CO}$ are isolated, in EtOH KOH , $(\text{PhNH})_2\text{CO}$, $(\text{PhNH})_2\text{CS}$ and PhNHCO_2N CHPh are the products. The β -oxime and II, heated to 120° , evolve COS and give $(\text{PhNH})_2\text{CO}$. In a sealed tube at room temp there results COS and $(\text{PhNH})_2\text{CS}$. In alk medium, the products are S , PhCN and $(\text{PhNH})_2\text{CS}$. I and *o*- $\text{MeC}_6\text{H}_4\text{NCS}$ in C_6H_6 at room temp give *o*- $\text{MeC}_6\text{H}_4\text{NH}_2\text{CS}$ (III) m 158° . In a sealed tube COS , III and PhCN were obtained, while in alk reaction the products are *carbo-o-toluido-o-benzaloxime*, m 124° , and *o*- $\text{MeC}_6\text{H}_4\text{NH}_2\text{CO}$, m 158° . I and *p*- $\text{MeC}_6\text{H}_4\text{NCS}$ behave similarly. MeCH NOH and II give COS and $(\text{PhNH})_2\text{CS}$ in a sealed tube or in an open vessel, in EtOH or C_6H_6 , autoxidation and decomposition take place. *o*- $\text{MeC}_6\text{H}_4\text{NCS}$ give S and III. In *o*- $\text{HOC}_6\text{H}_4\text{CH NOH}$ only the oxime group reacts with II, in alk soln the products are the nitrile and $(\text{PhNH})_2\text{CS}$, the possible reactions involved in the formation of the nitrile are discussed. *p*- $\text{MeOC}_6\text{H}_4\text{CH NOH}$ and II in EtOH at room temp give the nitrile and $(\text{PhNH})_2\text{CS}$, the β -isomer behaves in the same manner. Furfur *syn* aldoxime and II in EtOH - KOH give $(\text{PhNH})_2\text{CS}$, $(\text{PhNH})_2\text{CO}$ S and the nitrile. The *anti*-oxime behaves similarly in EtOH , while with alkali there results the carbanilino deriv, m 138° , identical with that obtained by the use of PhNCO . *o*- $\text{MeC}_6\text{H}_4\text{NCS}$ gives only *o*- $\text{MeC}_6\text{H}_4\text{NH}_2\text{CO}$. $\text{MeC}_6\text{H}_4\text{NCS}$ gives *carbo-o-toluidoacetoxime*, m $81\text{--}82^\circ$, the *p*-deriv m $105\text{--}6^\circ$. MeEtC NOH and II, without a solvent, give the carbanilino deriv, MeEtC NOCONHPh , m 135° , this also results in Me_2CO soln, in alk soln there is also formed $\text{H}_2\text{S}_2\text{O}_8$ or its decomp products H_2SO_4 and S . If the reacting compds are heated, there results the carbanilino deriv and $(\text{PhNH})_2\text{CS}$. *o*- $\text{MeC}_6\text{H}_4\text{NCS}$ gives *carbo-o-toluido-methyl Et ketoxime* m 80° the *p*-isomer m $146\text{--}7^\circ$. Carbanilino-diethyl ketoxime, m $96\text{--}8^\circ$ (25% yield) methylpropyl deriv, m 146° (13% yield) $\text{Me iso-Bu ketoxime}$ b 175° d $^{17}_D$ 0.8935 n $^{17}_D$ 1.456, carbanilino deriv, m 117° (20% yield); *carbo-o-toluido deriv*, m 184° . MePhC NOH gives a carbanilino deriv, m $131\text{--}5^\circ$, *carbo-o-toluido deriv*, m 109° *p*-isomer, ca 126° . Carbanilino-*acetophenoxime* pale yellow, m $118\text{--}20^\circ$. *o*- $\text{MeC}_6\text{H}_4\text{NCS}$ did not react. $\text{Ph}_2\text{C NOH}$ gives $\text{Ph}_2\text{C NOCONHPh}$, m 176° ($\text{Me}_2\text{NC}_6\text{H}_4$) $_2\text{C NOH}$ did not react with II or the *o*- Me deriv. Carbanilino-methyl *p*-tolyl ketoxime, m 112° . Monoximes of diketones do not react with II; α -

Isomeric acids, such as $\text{AcC}(\text{NOH})\text{CH}_2\text{CO}_2\text{H}$ or $\text{PhC}(\text{NOH})\text{OH}$, also do not react with II. C. J. WAT

The constitution of certain compounds formed by the action of alcoholic hydrochloric acid on unsaturated ketones. BATES CORRY AND HUGH RAN. *Proc Roy Irish Acad* B30, 471-9(1930), cf. C. A. 23, 3213, 24, 5293. The dimer formed by the action of alc. HCl on $\text{PhCH}=\text{CHCO}_2\text{Et}$ (I) yields a monoxime (II), m. 212-3°. It is reconverted into I by treatment with concd. H_2SO_4 or by boiling with alc. HCl and is unchanged by further treatment with NH_4OH (III), hence it is assumed that I is $\text{CH}_3\text{Ph}=\text{CHCH}(\text{CO}_2\text{Et})\text{CHCO}_2\text{Et}$ instead of the sym. $\text{EtCOCH}(\text{CHPh})\text{CH}(\text{CO}_2\text{Et})\text{C}$.

III. The compd. formed by treating II with P_2O_5 appears to be $\text{EtCOCH}(\text{CHPh})\text{CH}(\text{CHPh})\text{CHCONH}_2$ (IV). IV does not absorb H_2 nor react with III.

Heated with dil. H_2SO_4 , extd. with H_2O , PhH and CHCl_3 , neutralized with KOH, distd. and the distillate tested with I_2 in NaOH, IV gave the CHH test but none of the exts. left any residue. C. H. HAY

2-Hydroxy-4-methoxy- and 4-hydroxy-2-methoxybenzaldehyde. THOMPSON ELIZABETH DE KIRKPATRICK AND HENRY STEPHEN. *J. Chem. Soc.* 1931, 815-5. — $\text{MeO}-\text{C}_6\text{H}_4\text{CHO}$ (20 g.) in 70 cc. abs. H_2O and 20 g. HCN satd. with dry HCl and kept in a refrigerator for not more than 24 hrs. and the pptd. aldimine HCl dissolved in 2 l. cold H_2O and heated to boiling, gave 9.5 g. 2,4 $\text{HO}(\text{MeO})\text{C}_6\text{H}_3\text{CHO}$ (I) removed in a current of steam, and 10 g. of 4,2 $\text{HO}(\text{MeO})\text{C}_6\text{H}_3\text{CHO}$ (II). II phenylhydrazine, yellow, m. 134°, *p*-nitrophenylhydrazine, dark red, m. 255° (decompn.), semicarbazone, m. 212°, *p*-nitrophenylhydrazine, bright red, m. 256-5°, semicarbazone, lemon yellow, m. 217°, semicarbazone, m. 230°.

Condensation of aldehydes with hydrazones. II. Condensation of anisaldehyde and of salicylaldehyde with their respective phenylhydrazones. ANTONINO GIACALONE. *Gazz. chim. ital.* 60, 818-21(1930). — The expts. are in continuation of the previous ones (cf. C. A. 24, 837) and employ the same method. Each aldehyde can condense with its own phenylhydrazone or with the phenylhydrazone of another aldehyde, so that from $\text{H}_2\text{N}-\text{N}=\text{N}-\text{C}_6\text{H}_4-\text{CHO}$ (I), *o*- $\text{HO}-\text{C}_6\text{H}_3-\text{CHO}$ (II) and phenylhydrazine, 9 condensation products can be obtained. The earlier paper and present paper describe 3 of these, the remainder will be described in later papers. I (3 mols.), mixed with PhHNH_2 (III) (2 mols.), when the reaction is completed ZnCl_2 (10 g.) added, heated on a water bath until the mixt. becomes thoroughly hard, finely ground, washed with steam (to remove II), dried in the air, extd. with boiling C_6H_6 and the residue recryst. from boiling PhMe , yields *p,p'*-dimethoxybenzal-*p,p'*-dihydrazino-*p*-methoxytriphenylmethane, $\text{MeO}-\text{C}_6\text{H}_3-\text{CH}(\text{C}_6\text{H}_4\text{NHN}=\text{CHC}_6\text{H}_3\text{OMe})_2$, pale violet, m. 215-6°. Following the same procedure except that the hardened mass is washed with boiling EtOH instead of C_6H_6 , before recrystn. from boiling PhMe , II (3 mols.) and III (2 mols.) yield *o,o'*-dihydroxybenzal-*p,p'*-dihydrazino-*o*-hydroxytriphenylmethane, $\text{HO}-\text{C}_6\text{H}_3-\text{CH}(\text{C}_6\text{H}_4\text{NHN}=\text{CHC}_6\text{H}_3\text{OH})_2$, yellow, lvscent. 235°, m. 210°, sol. in dil. aq. NaOH from which solns. it is ripd by acids. III. Condensation of salicylaldehyde, of anisaldehyde and of *p*-nitrobenzaldehyde with benzalphenylhydrazine. *Ibid.* 821-3. — *o*- $\text{HO}-\text{C}_6\text{H}_3-\text{CHO}$ (5 g.) and $\text{PhCH}=\text{NNHPh}$ (I) (10 g.) treated with ZnCl_2 (5 g.) and a little PhMe to make a uniform paste, heated on a water bath, with agitation until the fused reaction mixt. becomes green and solidifies, a little water added, heated further, finely ground, washed with steam, dried in the air, dissolved in hot C_6H_6 , filtered and cooled, ppts. *p,p'*-dibenzaldehyde-*o*-hydroxytriphenylmethane, $\text{HO}-\text{C}_6\text{H}_3-\text{CH}(\text{C}_6\text{H}_4\text{NHN}=\text{CHPh})_2$ (II), m. 185°, insol. in dil. aq. NaOH (suggesting that the OH has taken part in the reaction, but this is disproved by the formation of an Ac deriv.). II, Ac_2O and fused AcONa refluxed 3 hrs., poured into water, let stand until solidified, dried, pulverized, dissolved in AcOH, poured into water and dried in air and then over H_2SO_4 in a desiccator, yield the Ac deriv., $\text{AcO}-\text{C}_6\text{H}_3-\text{CH}(\text{C}_6\text{H}_4\text{NHN}=\text{CHPh})_2$, m. 109-10°. Similarly *m*-nitrobenzaldehyde (5 g.), phenylhydrazine (11.5 g.) and ZnCl_2 (5 g.), heated until hard, steam distd., the residue washed with EtOH, dissolved in C_6H_6 and poured into ligroin, ppts. *p,p'*-dibenzaldehyde-*m*-nitrotriphenylmethane, $\text{MeO}-\text{C}_6\text{H}_3-\text{CH}(\text{C}_6\text{H}_4\text{NHN}=\text{CHPh})_2$, yellow, amorphous, m. around 125°, could not be crystd. from any solvent. *p*- $\text{O}_2\text{NC}_6\text{H}_3-\text{CHO}$ (2 g.), I (5.2 g.) and ZnCl_2 (2 g.), heated as before until hard, pulverized, boiled with water and then with EtOH, the residue purified by soln. in hot PhMe , yield *p,p'*-dibenzaldehyde-*p*-nitrotriphenylmethane, $\text{O}_2\text{NC}_6\text{H}_3-\text{CH}(\text{C}_6\text{H}_4\text{NHN}=\text{CHPh})_2$, orange yellow, m. 213-4°. C. C. DAVIS

A simple method for the preparation of phenylglyoxal. CARL NEUBERG AND EDUARD HOPMANN. *Biochem. Z.* 229, 413-5(1930). — The following reaction is utilized

for the prepn of phenylglyoxal $\text{PhCOCH}_2\text{NOH} + \text{ONO SO}_3\text{H} = \text{PhCOCHO} + \text{N}_2\text{O} + \text{H}_2\text{SO}_4$. Ten g of $\text{PhCOCH}_2\text{NOH}$ is dissolved in 35 cc. dioxane, dild with 10 cc. H_2O and cooled to 0° . With const. shaking 8.5 g. of $\text{NO}_2\text{SO}_3\text{H}$ crystals is added while the temp. is kept below 5° . Then the mixt. is warmed up to 40° , whereupon a vigorous reaction and evolution of gas occur. The flask must be kept in cold water to prevent the temp. from rising above 60° . This is now transferred with about 100 cc. H_2O to a Claisen flask and evapd to 70 cc. *in vacuo*. The only residue is extd 3 times with 50 cc. ether, and the combined ether exts. are washed several times with small quantities of H_2O . After removing the ether the PhCOCHO is extd. from the oil by boiling twice with 40-cc. quantities of H_2O . The hot soln. is filtered and evapd *in vacuo* to 30 cc., and on cooling with ice massive crystallization occurs. S. MOSCULIS

Dimethylallylbenzylacetic acid and isopropylbenzylvalerolactone. G. DASZKIN AND A. LÉVY. *Compt. rend.* 191, 1455-7 (1930). cf. C. A. 21, 581— $\text{Me}_2\text{C}(\text{CH}_2)_2\text{OH}$ with HBr gas gives $\text{Me}_2\text{C}(\text{CH}_2)_2\text{Br}$, b. 129° , which with $\text{PhCH}_2\text{CN}(\text{CO}_2\text{Et})_2$ gives $\text{Me}_2\text{C}(\text{CH}_2)_2\text{C}(\text{CH}_2\text{Ph})(\text{CO}_2\text{Et})_2$, b. $184-7^\circ$. This ester on sapon. and heating *in vacuo* to 160° gives 80% $\text{Me}_2\text{C}(\text{CH}_2)_2\text{C}(\text{CH}_2\text{Ph})\text{CO}_2\text{H}$, b. $157-60^\circ$, m. 49.5° , and 20% lactone, m. 25° and b. $148-50^\circ$. This does not condense to give a hydro-naphthalene deriv. V. F. HARRINGTON

Spectrochemical study of amino acid anhydrides. IV. Light absorption of derivatives of azlactones, diketopiperazine, hydantoin and thiohydantoin. TIZUCHI ASAHINA. *Bull. Chem. Soc. Japan* 5, 354-65 (1930); cf. C. A. 24, 298—The ultra-

violet absorption of azlactones of substituted hippuric acids, $\text{RCH}(\text{C}=\text{O})\text{O}(\text{C}_6\text{H}_5)\text{N}$ where R is a- (I), m- (II), or p- AcOC_6H_4 (III), o- (IV), m- (V), or p- MeOC_6H_4 (VI); of $\text{RCH}(\text{C}(\text{CO}_2\text{H})\text{NHCO}_2\text{R}')_2$ where R is furyl (VII), Ph (VIII), o-, m- or p- HOOC_6H_4 (IX), o-, m- or p- MeOC_6H_4 (XVI); dibenzal- (IX), and difuraldiketopiperazine (X), 4-benzal (XI), and 4-furalhydantoin (XII), 2-thiohydantoin and its following derivs. 3-acetyl-4-benzoyl, 3-acetyl-4-benzyl, 4-benzyl, 4-(p-hydroxybenzyl), 4-benzal (XIII) and 4-fural (XIV) VII, X, XII and XIV are more bathochromic and hyperchromic than VIII, IX, XI and XIII. I, II and III have an absorption max. near 3580 \AA . U, IV and VI have the same max. while V is less bathochromic. The azlactones are far more bathochromic than their hydrolysis products, XV and XVI. Absorption curves and methods of prepn. are given. V. F. HARRINGTON

Esterification of 3,5-diamino- and iodobenzoic acids with alcoholic hydrogen chloride. ANTON KARL AND ADOLF LARSENBERG. *Monatsh.* 56, 407-27 (1930)—The unimol. velocity coeffs. of esterification of 3,5-(H_2N) $_2\text{C}_6\text{H}_3\text{CO}_2\text{H}$ (I) and o-, m- and p- $\text{IC}_6\text{H}_4\text{CO}_2\text{H}$ with $\text{EtOH}\cdot\text{HCl}$ in the presence and absence of H_2O have been detd. at 25° . With EtOH contg. little H_2O the coeffs. for all the acids are proportional to the concn. of the HCl , in moist EtOH the coeffs. increase much more rapidly. The values for I show that the introduction of a 2nd m- NH_2 group into BrOH has a much smaller effect than the first. The retarding action of the halogen atom is most marked with the o-I acid and least with the m-isomer. For the o-halogenobenzoic acids, the coeff. decreases with increasing at. wt. of the halogen atom. All the m-halogenobenzoic acids have approx. the same coeff., which is less than that for BrOH . C. J. WEST

Behavior of some condensation products of p-dimethylaminobenzaldehyde toward bromine and nitrous acid. K. H. BAUER AND MARIE SEVPATH. *Ber.* 63B, 2691-5 (1930)—It had been shown that the condensation products of BrH with PhCH_2CN , $\text{NCC}_6\text{H}_4\text{CO}_2\text{Et}$ and other compds. do not add Br normally but are unreactive toward it. To det. whether this behavior is influenced by introduction of a Me_2N group into the Ph nucleus, the condensation products of p- $\text{Me}_2\text{NC}_6\text{H}_4\text{CHO}$ with PhCH_2CN , $\text{NCC}_6\text{H}_4\text{CO}_2\text{Et}$, $\text{CH}_3(\text{CN})$, MeCOPh and Me_2CO were treated with Br and HNO_2 . With p- $\text{Me}_2\text{NC}_6\text{H}_4\text{CH}(\text{C}_6\text{H}_5)\text{N}$ (I), p- $\text{Me}_2\text{NC}_6\text{H}_4\text{CH}(\text{C}(\text{CN})\text{CO}_2\text{Et})$ (II), p- $\text{Me}_2\text{NC}_6\text{H}_4\text{CH}(\text{C}(\text{CN})_2)$ (III), and p- $\text{Me}_2\text{NC}_6\text{H}_4\text{CH}(\text{CHCOPh})$ (IV), for the most part no Br or N_2O_3 is added at the double bond but Br is substituted in the $\text{Me}_2\text{NC}_6\text{H}_4$ nucleus in the o-position to the Me_2N group, as was shown by synthesis of the products, and N_2O_3 gives the corresponding NO_2 derivs. The yields in all cases were 60-75%. When, however, the PhCO group in IV is replaced by MeCO , i. e., with p- $\text{Me}_2\text{NC}_6\text{H}_4\text{CH}(\text{CHCOMe})$, there is formed, by primary addn. of Br at the double bond and subsequent elimination of HBr , α -p-dimethylaminophenyl- β -bromo- β -acetylstyrene (V), whose phenylhydrazone (VI) rearranges in AcOH on the H_2O bath into N-phenyl-3-methyl-5-p-dimethylaminophenylpyrazole (VII), with loss of H_2O . α -Phenyl-p-dimethyl-m-bromocinnamonsulfonate, from 3,4- $\text{Br}(\text{Me}_2\text{N})\text{C}_6\text{H}_3\text{CHO}$ and PhCH_2CN with a few drops KOH in alc. or from I with Br in CHCl_3 , green-yellow, m. 102° . m-Nitro analog, similarly prepd. from 3,4- $\text{O}_2\text{N}(\text{Me}_2\text{N})\text{C}_6\text{H}_3\text{CHO}$ and PhCH_2CN or from I with NaNO_2 in AcOH .

deguelin. Derric acid thus constitutes one-half of the mol of rotenone, deguelin and tephrosin. LAWRENCE P. MILLER

Rotenone, the effective constituent of derris root. VIII. The chemical constitution of derric acid and risic acid. SANICHI TAKEI, SHUICHI MIYAJIMA AND MINORI ONO. *Bull Inst Phys Chem Research (Tokyo)* 9, 1016-22 (1930), *Abstracts* 95-9 (in German) published with *Sci Papers Inst Phys Chem Research (Tokyo)* 15, Nos. 278-81; cf. *C. A.* 24, 4037.—Rotenone (20 g) and 50 g anhyd AcONa in 500 cc. EtOH , treated with 16 g I_2 in 130 cc. EtOH in 15 hrs, followed by boiling for an hr and filtering off the AcONa , gave 14 g of *dehydrorotenone* (I), m 216° , 4.3 g more is obtained from the EtOH filtrate. No Ac deriv of rotenone occurs. I (20 g) in 300 cc EtOH and 7 g KOH in 7 cc. H_2O give 16 g of *derric acid*, $\text{C}_{21}\text{H}_{14}\text{O}_7(\text{OMe})_2$ (II), m 153° , which with NH_4OH , HCl and AcONa gives 80% of *derric acid oxime*, $\text{C}_{21}\text{H}_{14}\text{O}_7\text{NOH}$, m. 142° . It gives with FeCl_3 the purple color, proving the presence of a phenolic OH radical. Ten g II in 200 cc. KOH (1%) with 75 g KMnO_4 in 700 cc. H_2O gives 2.6 g of *risic acid*, 2,5-(MeO) $_2$ $\text{C}_6\text{H}_3\text{C}(\text{OH})(\text{CO}_2\text{H})$ (III), m 225° , which is not changed by H_2O_2 oxidation in alk. soln., but entirely decompd by prolonged heating with KMnO_4 . III, methylated with CH_3N_3 gives quant. its *di Me ester*, $\text{C}_{21}\text{H}_{16}\text{O}_7(\text{OMe})_4$, colorless, m 80° , insol in NaHCO_3 but sol in NaOH ; with alkali it is converted into III. Three g III, heated for 7 mins. at $235-60^\circ$, gives 1.3 g of *decarbonylic acid*, $\text{C}_{21}\text{H}_{16}\text{O}_7(\text{OMe})_4$ (IV), m 116° , crystals from H_2O , m $68-69^\circ$ but m 116° when the water of crystal is lost. The oxidation of III and IV with N HNO_3 results in *dimethoxybenzoic acid* (V), yellow, m 209° , which with CH_3N_3 is converted quant. into the *Me ester*, $\text{C}_{21}\text{H}_{16}\text{O}_7\text{N}$, faint yellow, m 146° , being hydrolyzed by heating for 30 mins. with dil NH_4KOH to the original acid. The mixt. of 5 g III and 20 g KOH , moistened with a small portion of H_2O , heated for 3 hrs at $270-300^\circ$, gives 0.01 g of $\text{C}_{21}\text{H}_{16}\text{O}_7$, m $167-8^\circ$, whose mixt with hydroquinone m $168-9^\circ$. It reduces vigorously Fehling soln. and the NH_4AgNO_3 . One g III with 5 g KMnO_4 gives 0.1 g of *derric acid*, 2,5-(MeO) $_2$ $\text{C}_6\text{H}_3\text{C}(\text{OH})(\text{CO}_2\text{H})\text{CH}_2\text{CO}_2\text{H}$ (VI), colorless, m 171° , which, on fusion, does not liberate CO_2 and is oxidized to III. Regarding the optical isomers of VI and III, another paper will follow. K KONDA

Condensation products from arylldithioglycolic acids. ERICH GEBAUER FELNEGG AND HELENE JAROSCH. *Monatsh* 56, 317-21 (1930).—Brief details are given of dyes obtained by heating the following compds. with ClSO_3H at $50-60^\circ$ for 2 hrs: benzene-1,2, 1,3- (I) and 1,4-dithioglycolic acids, the 4-Cl and 2,5-Cl $_2$ derivs of I and 1,4-benzoquinone-2,3 and 2,5(2,6 2)-dithioglycolic acids. The dyes are probably bis(hydroxy)naphtheneindigo-x,x'-dithioglycolic acids. They are violet compds. which dye wool from acid baths blue-green tones. C J WEST

Synthesis of phenol ketones according to Hoesch. IV. W. ROSENTHAL AND K. DIACONT. *Ber* 63B, 2740-3 (1930), cf. *C. A.* 24, 93.—It was shown in Papers I and II that arylglyoxylic nitriles with multivalent phenols under the conditions of the Hoesch synthesis do not give hydroxylated benzils but lactone imines or ketones of polyhydroxytriphenylacetic acids, thus, PhCOCN and $m\text{-C}_6\text{H}_4(\text{OH})_2$ (I) give the lactone, $\text{HOC}_6\text{H}_4\text{CPh}(\text{C}_6\text{H}_4(\text{OH})_2)\text{COO}$ (II). This reaction seems, however, to be limited

to the arylglyoxylic nitriles and to depend on the linking of an aromatic acid radical with CN or CO , for on the one hand II has also been obtained (with NIEMANN) from PhCOCO_2H and I in Et_2O with $\text{ZnCl}_2\text{-HCl}$ and on the other hand the arylglyoxylic nitriles thus far tested (MeCOCN and EtCOCN) give with I under the conditions of the Hoesch synthesis only the monoketimines of the diketones 2,4-(HO) $_2\text{C}_6\text{H}_3\text{CO-COR}$. Phloroglucinol unexpectedly reacts much less readily and the products have such unpromising properties that nothing definite can as yet be said as to their compn. and structure. 1 (2,4-Dihydroxyphenyl)propane 1,2-dione, yellowish, m 150° , *disemicarbazone* yellowish, m $243-5^\circ$, *bis*(2,4-dinitrophenylhydrazones), red m $249-50^\circ$, 2 (2,4-Dihydroxyphenyl)butane 1,2-dione, yellowish, m 150° (the intermediate monoketimine HCl forms brownish crystals), *disemicarbazone*, m 205° , *bis*(2,4-dinitrophenylhydrazones), red, m 245° . C A R.

The condensation of catechol tannin. MAX BERGMANN AND GEORGI POJARLIEFF. *Naturwissenschaften* 18, 1114 (1930).—The condensation tendency of catechol is due to its structure of hydroxydihydropyran and by the simultaneous presence of a 2,3-double bond in the pyran ring with a hydroxyl in position 5. Tetramethylcatechol, with all phenol groups methylated, still is easily condensable (unpublished result). The benzene rings are not essential for the condensation, since glucal (Bergmann and Freudenberg, *C. A.* 24, 1352-3) is condensed in a similar manner by acids, dihydroglucal (a tetrahydropyran) is, however, acid resistant (Fischer, *C. A.* 8, 1121). Not all HO groups

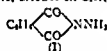
can be removed from catechol without destroying the condensability; anhydroaceto-butyl ale does not condense.

The chemistry of quinic acid. KARL JOSEPHSON *Scenik Farm Tils* 34, 649-54 (1930). —A critical review on the constitution of $C_8H_{12}(OH)_4CO_2H$. A. R. R.

Preparation of 4-nitrophthalimide and derivatives. LEONOLD F. LEVY AND HENRY STEPHEN *J. Chem. Soc.* 1931, 791-82. —Nitration of 20 g. $C_8H_4(CO)_2NH_2$ by adding it gradually to 25 cc. HNO_3 (d 1.44) and 200 cc. H_2SO_4 (10% SO_3) and holding the mixt. at 80° for 30 min., then cooling and pouring into 400 g. ice, keeping the temp. below 20° , gives 78% of the 4- NO_2 deriv., m. 108° , use of ordinary concd. H_2SO_4 reduces the yield to 40%. Reduction of 20 g. of the NO_2 deriv. with 81 g. $SnCl_2$ and 450 cc. HCl and 450 cc. H_2O gives 15.3 g. of the 4- NH_2 deriv., m. 224° , *Ac. deriv.*, pale straw, m. 331° (cor., decompn.). *Bz. deriv.*, cream colored, m. 322° (cor., decompn.). $C_8H_4(CO)_2O$ at 180° for 1 hr. gives 4 *phthalimidophthalimide*, yellow, m. 319° (cor., decompn.). The diazo compd. couples with $PhOH$ and α and β C_6H_5OH , giving orange, maroon and light brown dyes, resp., and with $PhNH_2$ and α and β $C_6H_5NH_2$ in acid soln., giving deep red, red and yellow dyes, resp. With $CuCl$ in HCl the diazo compd. gives the 4- Cl deriv., m. 210° , heating the diazo compd. gives the 4- H deriv., m. 281° .

C. J. WIST.

The action of hydrazine acetate on 3-nitrophthalic anhydride. M. MIHAILESCU AND L. PROTODRASCUL *Bul. soc. chim. Roum.* 12, 95 [92 (1930)]. —M and II reply to the objections of Radulescu and Georgescu (cf. C. 1 20, 181) to the formula (I) which Mihailescu and Protodrascul (cf. C. A. 18, 3143) propose for the product of reaction of $N_2H_4 \cdot 2AcOH$ on $C_8H_4(CO)_2O$ rather than the formula (II) of Forsterling.



M and II cite the following properties of this compd. which are not consistent with formula II and its enol (III) but are consistent with formula I and its hydrolysis product $HO_2CC_8H_4CONHNH_2$ (IV). (1) The substance is difficultly sol. in water but is hydrolyzed by water to give an acid reaction to litmus. (2) It dissolves completely in alkali hydroxides with red color and is pptd. unchanged by acids. (3) It dissolves in carbonates and bicarbonates with effervescence but it is not pptd. by an excess of CO_2 , as R. and G. erroneously state. (4) Only monometallic salts are formed. (5) Three properties indicate that the substance possesses an NH_2 group: (a) It is decompd. by HNO_3 and $NaBrO$. (b) It reacts with aromatic aldehydes to give condensation products; (c) its *Ac* and *Me* derivs. have properties more like those of *N*- than of *C* derivs. Ten g. of 3-nitrophthalic anhydride in 150 cc. boiling glacial *AcOH* reacts with 15 g. $N_2H_4 \cdot H_2SO_4$ and 18 g. cryst. $AcONa$ in H_2O to form $NO_2C_8H_4(CO)_2NNH_2$ (V). V does not m. 300° ; $NaBrO$ decomp. it, only mono-*Ac* derivs. are formed by Ac_2O and $AcCl$, alkali hydroxides, carbonates and bicarbonates dissolve it, producing a red coloration and needle-shaped, golden yellow crystals of only monometallic derivs. which do not m. 300° but above this temp. decomp. with explosion. Ac_2O reacts with V to form the mono-*Ac* deriv. as needles resembling cotton wool which m. 230° . V reacts with $BzCl$ in hot pyridine to form a mono-*Bz* deriv. as a sandy powder m. 210° . V in hot NH_4OH reacts with $AgNO_3$ to form a deep yellow *Ag* salt which is a monometallic salt as indicated from N. elemts. *Ag* could not be detd. as explosions resulted in the attempts. The *Ba* salt of V was prepd. by decompn. of the *K* salt of V with $BaCl_2$ as yellow needles which analysis showed to be a monometallic salt. V reacts with Me_2SO_4 at 120° to form plates of a mono-*Me* deriv. m. 265° . IIating V with anisaldehyde and fused $AcONa$ gives needles of $NO_2C_8H_4(CO)_2NNCH_2CH_2OMe$ (VI) which m. 280° and react with KOH soln. to form $MeOC_8H_4CHO$ and the *K* salt of V. These properties of V together with the arguments in regard to the structure of I (an analogous compd.) justify the structure assigned to V. OREN E. SHEPPARD.

Menthone series. DL. New optical resolution of *dl*-menthol and of *dl*-camphor-10-sulfonic acid, JOHN READ AND WM. J. GRUBB *J. Chem. Soc.* 1931, 188-95, cf. C. A. 25, 1233. —Although the naturally occurring form, *l* menthol (I), was fairly readily obtained by the optical resolution of the *dl* form (II) by Pickard and Littlebury (C. A. 6, 1284), the method does not appear to lend itself to the prepn. of *d* menthol (III) in quantity. *l* and *d* camphor-10-sulfonyl chloride (IV) in quinoline or pyridine give 85% of *l* menthyl *d*-camphor-10-sulfonate (V), m. 125.5° , crystallizes from light petroleum in slender needles, 5-6 cm. long, from $AcOEt$ in doubly-terminated prisms; the system is tetragonal, $c/a = 1.3788:1$, $a(100)$, $m(110)$, $p(111)$; $[\alpha]_D^{25} -20.9^\circ$ ($CHCl_3$,

c 19353). At 153°, V decomps. to *d*-camphor 10-sulfonic acid (VI) and partly racemized *d* Δ¹ menthene (VII). When steam-distd., V undergoes extremely slow decomposition, giving I, VI and VII. V is very stable toward hot FIOH KOH, Na in hot FIOH gives I and VII. Prolonged boiling with 20% HCl or H₂SO₄ causes a gradual decomposition of the ester, the most satisfactory hydrolyzing agent is hot 50% (XXIII), the products being I, VI and VII. *2-Methyl-1-camphor-10-sulfonate* (VII), m 47°, $[\alpha]_D^{25} -77.6^\circ$ (CHCl₃, c 1.9300), at 153° it decomps. into *l*-camphor 10-sulfonic acid and VII. It is somewhat less resistant to hydrolysis than is V. II and IV in quinoline give an ester with $[\alpha]_D^{25} 2^\circ 0'$ (CHCl₃, c 1.0), 2 crystals from light petroleum and 2 from AcOFt give 45% of V; decomposition with 50% (CO₂H)₄ gives only 10% of pure I. II and *l*-camphor 10-sulfonate give 42% of *d* methyl-1-camphor 10-sulfonate (IX), m 121.5°, $[\alpha]_D^{25} 21.9^\circ$ (CHCl₃, c 2.0740), decomposition with 50% (CO₂H)₄ gives 10% pure III. An equimol. mixt. of VIII and IX melted indefinitely over a wide range of temp., and gave no indication of the formation of a partial racemate. Similar behavior was shown by a mixt. of IV and VIII. In both cases, the observed rotatory power in CHCl₃ was practically equal to the mean value of the 2 components. An equimol. mixt. of IV and IX appears to give mixed crystals when allowed to sep. slowly from AcOFt. Details are given of the resolution of *d*-camphor 10-sulfonic acid with L.

C. J. West

Action of Beckmann's mixture on some monocyclic terpenes. II. Terpinolene and "organene." THOMAS A. HENRY and HUMPHREY PAGET. *J. Chem. Soc.* 1931, 25-32, cf. C. A. 22, 1345.—Terpinolene (I) was prepd. by shaking *l*-pinene with half its vol. of EtOH H₂SO₄ (50% by wt.) and also from terpenes by boiling gently with 40% aq. (CO₂H)₄ for 30 min. or by keeping it at room temp. for 2-3 days with EtOH H₂SO₄. I was purified through the tetrabromide, which crystallizes in 2 forms, from 2 vols. EtOH, MeCO or AcOEt, m 122°, monoclinic, $a \times b \times c = 0.7804 \times 1.0768 \times 1.14^\circ 15'$ (cf. Hintze, *Z. Krist.* 10, 258 (1933)), from 10 vols. MeCO monoclinic, rectangular plates, m 119°, $a \times b \times c = 2.5905 \times 1.2033 \times 1.11^\circ 8.5'$, $d(100) = (110)$, $d(001)$, $d(101)$. The regenerated I has 121°, b.p. 140°, $d_4^{20} 0.8502$, $n_D^{20} 1.4866$ the same product is obtained from the 2 tetrabromides and on bromination yields 80-85% of the first tetrabromide. I is not very readily attacked by Beckmann's mixt. The principal products of oxidation are levulinic acid and AcOH. A small amt. of an unsatd. lactone is obtained, which is regarded as 3-ethoxy-2-hydroxy-1-(3-oxopropyl)-Δ¹-keto-Δ²-lactone (II), C₁₀H₁₆O₄, m 42-50°, semicrystalline, m 220-3° (dec., decomps.), reduction of II gives homoterpenyl Me ketone. Hydrolysis with warm alkali gives the acid AcCH₂CH(CMe)₂CHCO₂H, m 130°, which absorbs 4 H, giving *d*-ε-keto-δ-isopropylheptonic acid. II is not readily attacked by Beckmann's mixt., but at 60° after 2 days there results terebithic acid in quant. yield. "Organene" (Pekles, C. A. 2, 2248, 2673) is generally regarded as α-terpinene; this suggestion ignores the fact that organene yields a nitrovinyl chloride, m 91-4° a nitropropene, m 180°, and a nitrobenzylamine, m 104-5°, which have not been obtained from α-terpinene. Recrystallization of residues from the prepn. of carvadiol from *l*-pinene organene oil shows that α-terpinene is present, but there is also about 14% of a terpene giving the above cryst. derivs. and for which the name organene is again proposed until the constitution can be established.

C. J. West

Pinene and nopinene. GEORGES BRU. *Bull. Inst. chim.* 1929, 271-6, 293-300, 1930, 7-16, 33-8, 63-8, 87-94, 104-12, 131-4, 151-4, 175-8, 194-203, 225-8, 244-7, 270-4 cf. C. A. 17, 317; 19, 447-5, 55; 22, 154; 3157; 23, 829.—After reviewing the origin and phys. and chem. properties of *d* and *l*-pinene and *d*- and *l*-nopinene, a more detailed account is given of the previously published investigations. A. P. C.

Production of camphor. WILLIBALD REHMEN. *Rivista Ind.* 5, 105-8 (1930).—The present and past methods for the distn. of camphor producing material are described.

W. O. E.

Photopolymerization of p-menthene. W. TALLIS. *Sci.* 63B, 2758-40 (1930).—Piperitone I in aq. AcOH or alc. exposed to the light of a quartz Hg lamp begins after several hrs. depending on the diam. of the containing quartz vessels) to deposit needles which are repeatedly filtered off until no more are formed, they are followed by the sepn. of a yellow oil. This oil, when dissolved in hot alc., gives more of the needles on cooling. The yield of cryst. product is the same whether the soln. is illuminated in an air or H₂ atm. and small quantities of uranyl acetate have but a slight influence if the deposits formed on the walls of the vessel are removed, larger quantities, through their screening action, slow up the reaction. The needles seemed to be quite homogeneous, b. about 365° almost without decomps. and m. 112°, but by re-

peated crystals from MeOH they were sepd into 3 fractions having the compn and mol wt (Rast) of a dimeric piperidine: (1) needles, m 162° , form no semicarbazone, (2) silky needles, m $142-4^{\circ}$, yielding a semicarbazone, $C_{10}H_{16}O_2N_4$, m $228-35^{\circ}$ (decompn) depending on the rate of heating, that sepd in 2 forms (needles and cubes) which did not depress each other's m p. (3) rhombic leaflets, m $157-0^{\circ}$. (1) and (3) greatly depressed the m p of each other. The semicarbazone of (2) regenerates (2) with hot concd aq $(CO_2H)_2$. The dimerization in this case evidently has occurred through the double bond in I, probably with formation of a 4 membered ring, but the non-formation of a semicarbazone by (1) and (3) does not necessarily prove that the C O group is involved in the formation of these dimers, for steric influences may prevent the formation of a semicarbazone even though the C O group is still intact. In sun light instead of the light of the Hg lamp I yields only the dimer (1). Preliminary expts indicate that carvone which, according to Ciamician and Silber, yields only 1 dimer in sunlight, also forms 2 dimers when exposed to the light of a Hg lamp.

C. A. R.

Correction [autoxidation of cedrene] A. BILMANN, *Ber* 63B, 2880, 1930), cf. C. A. 24, 97. The unsatd terpene $C_{15}H_{24}$ (obtained by dehydration of the alc resulting from the oxidation of cedrene with O and a Co catalyst) was described as having $[\alpha]_D^{25} = -125^{\circ}$ when, as a matter of fact it is d rotatory. The prepn used at the time was insufficiently purified and final values on the constants of the compd will be published shortly.

C. A. R.

Fokienol, a new monocyclic sesquiterpene alcohol. L. S. GLITCHIK, *Comp. rend* 191, 1437-60, (1930). Distn of stems of *Fokienia kadsura* gives an oil which is mostly a tertiary monocyclic alc, $C_{15}H_{26}O$, b $_p$ 125° c, $d_4^{20} 0.9236$, $n_D^{20} 1.4975$, $n_D^{25} 1.4835$, mol refraction 70.18, indicating 2 double bonds. Distn causes ring closing and the compd must be isolated through its formate. Acids cause dehydration to a bicyclic diene, $C_{15}H_{24}$, isodakienene. HIO_3 gives the monocyclic lakienene, b $_p$ $112-4^{\circ}$, $d_4^{20} 0.8802$, $n_D^{20} +16^{\circ}30'$, $n_D^{25} 1.49394$. All 3 on dehydrogenation with S give 1,6,4- $C_{15}H_{14}MeClIMe$.

V. I. HARRINGTON

Heteropolar carbon compounds. XII. New dyestuffs of the aniline blue series and perchlorates of several important triphenylmethane dyes. W. DILTMAN AND R. DINKLAGE, *J. prakt. Chem* 129, 21-30, (1931), cf. C. A. 24, 1855.—Parafuchon base and $p-H_2NC_6H_4Ph$, heated with H_2O 6 hrs. at $175-80^{\circ}$, give *tri(4-phenyl)aniline blue*, $(PhC_6H_4NHC_6H_4)_3CCl$, golden glistening crystals with a violet blue base color, decomp about 225° ; drying at 110° causes no noticeable decompn. Fuchon base and $p-H_2NC_6H_4OPh$ in the same way give *tri(4-phenyl)aniline blue*, $(PhOC_6H_4NHC_6H_4)_3CCl$, violet, decomp. about 230° . $p-H_2NC_6H_4SiPh$ and parafuchon base give *tri(4-phenyl)aniline blue*, $(PhSC_6H_4NHC_6H_4)_3CCl$, blue or violet, decomp. $220-31^{\circ}$. Colors are given for the H_2O , H_2O and H_2SO_4 solns. Little difference is observed in the color of the 3 new dyes. The m. ps. for the following perchlorates are approx. Benzaldehyde green, blue-green, 20° ; brilliant green, blue-green, 143° ; rhodamine 3B, red, 167° ; crystal violet, violet, $200-40^{\circ}$; methylene blue, dark blue, $231-2^{\circ}$; Fuchsin blue, blue, in a bath at 320° , decomp. violently; pararosaniline, dark red, 317° (block).

C. J. WEST

Conjugated double bonds. XIV. Addition of maleic acid anhydride to polyenes. RICHARD KATZ AND THEODOR WAGNER-JUREG, *Ber* 63B, 2862-79, (1930); cf. C. A. 25, 1083.—The addn. of maleic anhydride (I) to dienes has often been used in the study of the structure of natural products such as cholesterol and muscarulin, and it seemed of theoretical and practical importance to det. whether the reaction could be extended to compounds with a large no. of conjugated double bonds such as the dyes of the buxin and carotene series. Preliminary expts. showed that all the polyene dyes tested are decolorized by I and an addn. product of Me buxin was isolated as a well-crystd. Na salt. Because of the no. of isomers possible, however, the structure of this product could not for the time being be established and it was to shed light on this question that the present study of the behavior of a series of synthetic diphenylpolyenes toward I was undertaken. Stilbene, although it has only 1 double bond, reacts very readily with I in hot xylene, forming an amorphous, approx. 1:1 addn. product of high mol. wt.; it is presumably cyclic and is to be described in detail elsewhere. $(PhCH_2CH=CH)_2$ takes up 1 mol. I, as has also since been found by Diels and Adler (C. A. 24, 96) but the addn. products vary with the conditions of prepn. In xylene are obtained rhombs m 198° to 203° from prepn. to prepn., which can be sepd. without change in the compn., $C_{24}H_{20}O_4$, into fractions differing in m. p. by at most 4° and which are probably mixts. of isomers. D and A, carried out the addn. in a molten mixt. of the components and

obtained a compd m 207° (apparently uncor). An isomer m 212.5-4.5° (cor) was obtained by rearrangement of the above mixt with $\text{Cl}_2\text{CHCO}_2\text{H}$ in CHCl_3 , but its m p fell to 196-8° (cor) on crystn from C_6H_6 . That the 1 has added at the 1,4 positions in every case is shown by the conversion of the addn products into terphenyl by hydrolysis and oxidation with $\text{K}_2\text{Fe}(\text{CN})_6$. The isomers probably differ in the position of the double bond. The isomer m 198° with hot aq NaOH gives a 3,6-diphenyltetrahydrophthalic acid (II), rhombs with 1 H_2O from alc., which when heated loses H_2O and shows the high m p of D and A's anhydride. An oily isomer (III) of II is obtained through the dichloride, m 143-4° (cor), from $(\text{PhCH CH})_2$ and fumaryl chloride. With HCl in Et_2O III changes into crystals, m 230-1° (cor). From its method of prepn, III contains the CO_2H groups in the *trans* position. Assuming that the double bond occupies the same position in II and III, *cis trans* isomerism can be explained only if the double bond is in the 3,4- or 4,5-position. The low melting anhydride of II decomps to a large extent, when distd under 12-5 mm, into $(\text{PhCH CH})_2$ and I. $(\text{PhCH CHCH CH})_2$ with 1 yields an addn product $\text{C}_{24}\text{H}_{16}\text{O}_4$ (IV), m 199.5-200° (cor), confirming the observation of D and A. That addn takes place at the 1,6-positions is shown by simultaneous decarboxylation and dehydrogenation by distn with anhyd $\text{Ba}(\text{OH})_2$ and Se, which yields 1-phenyl 2-biphenylstyrene (V), m 221-2° (cor), identical with the product obtained by distn of $p\text{-PhCH}_2\text{COCH}_2\text{Ph}$ with Zn dust in H_2 . $(\text{PhCH CHCH CH})_2$ likewise adds 1 mol I in boiling xylene but the product (VI) is not identical with the 3-phenyl-6-styryltetrahydrophthalic anhydride, (IV) above. It m 192-3° (cor) and is also obtained by distn of IV *in vacuo*; it is probably also a 3-phenyl 6-styryltetrahydrophthalic anhydride, for with O_2 it yields BrOH in abundance, and hydrolysis gives an acid which, when dehydrogenated with $\text{K}_2\text{Fe}(\text{CN})_6$ and decarboxylated or when distd directly with soda lime, changes into V. The isomerism of IV and VI is probably due to a difference in the position of the double bond in the cyclohexene ring. On hydrolysis VI gives an acid m 190-1° (cor) with regeneration of VL $(\text{PhCH CHCH CH})_2$ adds 2 mols I when the mixt of the components is cautiously fused, forming 6,6'-bis(3-phenylcyclohexene-1,2-dicarboxylic anhydride) (VII), m 275° (cor, decompn), and in boiling xylene is obtained an isomer (VIII), m 264° (decompn), along with VII. Crystn from $(\text{CH}_3\text{CHCO})_2\text{O}$ raises the m p of VIII to 292° (decompn) which again falls to 264° on crystn from Ac_2O or $p\text{-C}_6\text{H}_4\text{Cl}_2$. On hydrolysis, dehydrogenation with $\text{K}_2\text{Fe}(\text{CN})_6$, and distn with soda lime, VIII gives $(p\text{-PhC}_6\text{H}_4)_2$, m 313-4° (cor), attempts to effect the degradation by direct distn with soda lime gave a hydrocarbon, $\text{C}_{24}\text{H}_{16}$, m 281-2° whose structure has not been established. Hydrolysis of VIII with concd NaOH gives an amorphous dicarboxylic acid, easily sol in cold HCO_2H and soon sepg in crystals probably as the result of isomerization, the same change was observed when the amorphous acid was warmed in Et_2O with a few drops of HCO_2H . The isomer seps from dil alc with 2 mols H_2O and when heated to 140° in $\text{Cl}_2\text{CHCO}_2\text{H}$ and pptd with Et_2O it gives the anhydride m 290° (see above). Attempts to condense $(\text{PhCH CHCH CH})_2$ with only 1 mol I gave always VII and the unchanged hydrocarbon. Toward fumaryl chloride the tetraene behaves as with I; the resulting tetracarboxyl chloride m 231.5-25° (cor) $(\text{PhCH CHCH CH})_2$, in boiling xylene adds chiefly 2 mols I; there are obtained 2 isomers, $\text{C}_{24}\text{H}_{16}\text{O}_4$, sepd by fractional crystn from Ac_2O into the more sol chief product, m 235.5°, and a less sol isomer, m 255.5° (both cor, decompn). There is also formed an amorphous substance (apparently a 1:3 product), decompn 276.5° (cor). The 235.5° compd is apparently formed by 1,4 and 7,10-addn, dehydrogenation and decarboxylation gives a compd m 302-3°, apparently identical with a product obtained by dehydrogenation of $(p\text{-PhC}_6\text{H}_4\text{CH}_2)_2$ and assumed to be di-*p*-biphenylstyrene. $(\text{PhCH CHCH CHCH CH})_2$ reacts readily with I in boiling tetralin, forming a product difficultly sol even in boiling Ac_2O , dioxane, PhNO_2 , and BrOEt . Apparently 3 mols I are in the main added. Simultaneous dehydrogenation and decarboxylation gives a hydrocarbon, begins to m 320°, having the compn of a quingnaphenyl. $(\text{PhCH CHCH CHCH CHCH CH})_2$ behaves like the hexaene, yielding a difficultly sol product which has not yet been obtained in homogeneous form. These results show that when a large no. of conjugated double bonds are present they each add 1 mol I with formation of 6 membered rings. The reactive diene groups in the diphenylpolyenes are at the ends of the chains and the behavior of these compds toward I is to be ascribed to the same cause as their behavior toward H in the formation of dibenzylpolyenes, viz, the greater reactivity of the end methine groups adjacent to the Ph residues.

C A R

2-Methylnaphthalenes. III. K. DZIEWOŃSKI AND A. WULFSOHN. *Bull intern acad polonaise* 1929A, 143-8.—Sulfonation of 2- $\text{C}_{10}\text{H}_7\text{Me}$ with an equimol quantity

of ClSO_3H in PhNO_2 at $30-40^\circ$ gave 2-methylnaphthalene-8-sulfonic acid (Na salt; chloride, m $94-6^\circ$, amide, m $172-4^\circ$, anilide, m $162-4^\circ$). Its constitution was established by conversion into 7-methyl- α -naphthol, m $109-11^\circ$ (acetate, m $39-41^\circ$, *azo deriv* with *p*-nitroaniline, m 250° (decompn)), by fusion of the Na salt with KOH.

B. C. A.

1,4-Dibenzyl-naphthalene and its corresponding keto derivatives. K. DZIEWOŃSKI, J. MOSZEW, S. LEPIANEKOWICZ AND L. SECHER. *Bull intern acad polonaise* 1929A, 650-7. —The dibenzyl-naphthalene, m 88° obtained as one of the products of the action of PhCH_2Cl on C_{10}H_8 in presence of AlCl_3 or fused ZnCl_2 (β -dibenzyl-naphthalene" of Dzewonski and Moszew, *C. A.* 23, 3220, 3921) is now shown to be 1,4-dibenzyl-naphthalene (picrate, m 107°), also prep'd by the reduction of 4-benzoyl-1-benzyl-naphthalene, m 113° (phenylhydrazone, m $155-6^\circ$, oxime, m $202-3^\circ$). The latter is obtained by the interaction of 1- $\text{C}_{10}\text{H}_7(\text{CH}_2\text{Ph})$ and Br_2Cl in presence of fused ZnCl_2 . The orientation of both compds is established by their oxidation with HNO_3 to the known 1,4- $\text{C}_{10}\text{H}_6\text{Br}_2$, m 100° (Scholl and Neumann, *C. A.* 16, 2143) (dioxime, m 261°). Treatment of 1,4- $\text{C}_{10}\text{H}_7(\text{CH}_2\text{Ph})_2$ with ClSO_3H at the ordinary temp affords 1,4-dibenzyl-naphthalene-sulfonic acid isolated as the Na salt. Nitration of 4,1- $\text{C}_{10}\text{H}_6\text{Br}_2$ yields a mononitro deriv, m $172-3^\circ$. A by-product of the reduction of the same compd is 1,4-dibenzyltetrahydronaphthalene m $92-3^\circ$.

B. C. A.

peri-Benzoyl and -benzyl derivatives of naphthalene. K. DZIEWOŃSKI, J. AVERBACH AND J. MOSZEW. *Bull intern acad polonaise* 1929A, 658-63. —1-Benzoyl-8-benzyl-naphthalene, m 142° (oxime, m $220-2^\circ$), is prep'd by the action of PhCH_2Cl on 1- $\text{C}_{10}\text{H}_7\text{Br}$ in the presence of AlCl_3 at $160-70^\circ$. Oxidation of this with HNO_3 yields 1,8- $\text{C}_{10}\text{H}_6\text{Br}_2$ (bisphenylhydrazone, m $270-1^\circ$, dioxime, m 270°) (cf *C. A.* 23, 3220, 3923).

B. C. A.

Synthesis of acetyl derivatives of 1-benzyl-naphthalene. I. 4-Acetyl-1-benzyl-naphthalene. K. DZIEWOŃSKI AND J. MOSZEW. *Bull intern acad polonaise* 1930A, 66-71. —1- $\text{C}_{10}\text{H}_7\text{CH}_2\text{Ph}$ reacts with AcCl in the presence of AlCl_3 in CS_2 at the ordinary temp to give 4-acetyl-1-benzyl-naphthalene, b₁₁ $240-5^\circ$, m 75° (picrate, m 113°) (60%), together with a di-*Ac deriv*, b₁₁ $265-70^\circ$, m 135° (10%), as a by-product. The oxime, m $240-1^\circ$, of 1,4- $\text{C}_{10}\text{H}_6(\text{CH}_2\text{Ph})_2\text{Ac}$ undergoes a Beckmann rearrangement when its soln. in AcOH and Ac_2O sat'd with anhyd HCl is heated at 100° , to give 4-acetamido-1-benzyl-naphthalene, m $208-9^\circ$, hydrolyzed by boiling 6% HCl to 4,1- $\text{C}_{10}\text{H}_6(\text{CH}_2\text{Ph})\text{NH}_2$, identical with a specimen obtained by the action of NH_3 on 4,1- $\text{C}_{10}\text{H}_6(\text{CH}_2\text{Ph})\text{O}_2\text{H}$. (*C. A.* 22, 2164). HNO_3 (d. 1.52) in AcOH at $10-15^\circ$ converts 1,4- $\text{C}_{10}\text{H}_6(\text{CH}_2\text{Ph})_2\text{Ac}$ into its 5-nitro-deriv, m 153° , while boiling 10% HNO_3 oxidizes it to 4-benzoyl-naphthoic acid, m $180-1^\circ$.

B. C. A.

Bromination of 4-nitro-1-methylnaphthalene. JOHN S. H. DAVIES AND ALBERT E. OXFORD. *J. Chem. Soc.* 1931, 230-1. —4,1- $\text{O}_2\text{NC}_{10}\text{H}_7\text{Me}$ and Br in a quartz flask in ultra-violet light give a mono-*Br deriv*, yellow, m $137-5-8-5^\circ$, and a very small quantity of a tetra-*Br deriv*, m $189-90^\circ$ (decompn), the same mixt was obtained in a sealed tube at $100-20^\circ$ for 2 hrs. α - $\text{C}_{10}\text{H}_7\text{CH}_2\text{Cl}$, b₁₁ $146-53^\circ$, seps. from EtOH at -15° as silvery flakes, m $29-5-30-5^\circ$, this was previously described as a liquid. 2,4- $(\text{O}_2\text{N})_2\text{C}_6\text{H}_3\text{CH}_2\text{Cl}$ and NaBr in MeOH , boiled 6 hrs, give the bromide, m $46-7^\circ$.

C. J. WEST

Naphthalene series. II. Diaryl- and triarylmethane derivatives of dimethyl- α -naphthylamine. BASKER GOKHLE AND FREDERICK A. MASON. *J. Chem. Soc.* 1931, 118-26. cf. *C. A.* 24, 5296. — $(\text{Me}_2\text{NC}_{10}\text{H}_7)_2\text{CH}_2$, m $181-2-5^\circ$, was obtained from α - $\text{C}_{10}\text{H}_7\text{NMe}_2$ and HClHO in AcOH - HCl and methylal in AcOH - HCl or H_2SO_4 , the base was surprisingly resistant to oxidizing agents; chloranil in glacial AcOH at 110° gives 2% of 4,4'-tetramethyldiamino-1,1'-dinaphthylcarbinol, m $184-6^\circ$; cold AcOH gives a faintly green soln, becoming blue-violet on warming; concd. HCl or H_2SO_4 gives a deep yellow soln. 4-Dimethylamino-1-naphthyl chloride and anilide were prep'd from the acid. Attempts to prep the ketone $(\text{Me}_2\text{NC}_{10}\text{H}_7)_2\text{CO}$ from this chloride and α - $\text{C}_{10}\text{H}_7\text{NMe}_2$ in $\text{C}_2\text{H}_5\text{Cl}$ gave only tarry products. 4,4',4"-Hexamethyltrisaminotrisnaphthylmethane, m $266-7-5^\circ$, is formed from α - $\text{C}_{10}\text{H}_7\text{NMe}_2$ and $\text{HC}(\text{OEt})_3$ with AlCl_3 at 70° . $\text{Me}_2\text{NC}_{10}\text{H}_7\text{COCl}$ and PhNMe_2 with AlCl_3 in $\text{C}_2\text{H}_5\text{Cl}$ gave 4,4'-tetramethyldiaminophenyl-1-naphthyl ketone, m $128-5-9^\circ$, Et_2O , EtOH and dil. mineral acids give light greenish yellow solns. With POCl_3 the ketone condenses with PhNMe_2 , yielding a deep blue dye, probably identical with "Naphthoblauf". Reduction with Na-Hg in EtOH gives the corresponding carbinol, pale yellow, m $62-3^\circ$; the AcOH soln is a brilliant blue. 4- $\text{H}_2\text{NC}_{10}\text{H}_7\text{NMe}_2$ obtained by reduction of the 4-NO deriv. yields an *Ac deriv*, m 185° . Attempts to prep. the analog of methylene-blue B failed. Ab-

sorption curves of the carbinols in AcOH are given, values calculated according to Moir's theory agree with the observed maxima

C. J. WEST

Naphthalene-1,5-disulfonic acid as a by-product in the monosulfonation of naphthalene. CYRIL H. RADCLIFFE and WALLACE P. SHORT. *J. Chem. Soc.* 1931, 220. In the prepn of $C_{10}H_7CN$ from $C_{10}H_7SO_3Na$ and KCN, there is obtained 1.4 g. 1,5- $C_{10}H_6(SO_3)_2$ from 100 g. $C_{10}H_7$ employed in the sulfonation, because of the sparing soly of $C_{10}H_7(SO_3H)$ as the Na salt and the incomplete conversion into the dicyanide, this est. of the degree of disulfonation is a conservative one

C. J. WEST

Theory of halogen substitution II. PAUL PRIFFER and PAUL SCHNEIDER. *J. prakt. Chem.* 129, 129-44 (1931), cf. *C. A.* 24, 1858. Previous work showed that certain ethylenes give colored bromides. In order to study the constitution of these the following work was carried out. $Ph_2C=CH_2$ and $(p-MeC_6H_4)_2C=CH_2$ gave absolutely no color with Br. $\alpha-C_{10}H_7MgBr$ and $AcOEt$ give α,α -dinaphthylmethylcarbinol, m. 146°, giving an intense malachite green color with coned H_2SO_4 and a brick red color with Br. With coned HCl there results α,α -dinaphthylethylene, m. 107°, coned H_2SO_4 gives the same color as with the carbinol, the ethylene does not show a characteristic behavior with Br, in CCl_4 Br gives a mono-Br deriv., m. 148°, excess of Br gives a tetra-Br deriv., m. 69-70°, sol in 30% oleum with a grass-green color. Attempts to prep. $(C_{10}H_7)_2C=CHPh$ by the action of $PhCH_2CO_2Et$ upon $\alpha-C_{10}H_7MgBr$ gave only $PhCH_2COCHPhCO_2Et$. Biphenylphenylethylene, $PhC_6H_4CPh=CH_2$, m. 91-5°, coned H_2SO_4 gives a brick red color, Br vapor colors the compd. greenish gray, Br deriv., m. 162°, warming with coned H_2SO_4 gives an intense red-violet color, Br vapor colors the compd. black-violet and the color is very stable, the di-Br deriv. m. 155°, Br vapor does not color the crystals. $(PhC_6H_4)_2C=CH_2$, m. 211°, gives a Br deriv., m. 187-8°, colored by Br vapor a red brown, which color quickly disappears, di-Br deriv., m. 194-5° asym. Diphenylmethylstyrene, $(PhC_6H_4)_2C=CHPh$, light yellow, m. 165-6°, sol in coned H_2SO_4 with a deep red violet color, Br vapor gives a deep green color stable for only 1-2 sec. Br deriv., m. 150-2°, Br vapor gives a dirty grayish black color, stable for about 1 hr. asym. Diphenylphenylethylene, m. 192-3°, difficultly sol in cold coned H_2SO_4 and giving a deep bluish red color on warming, Br vapor gives a grass-green color, stable for only 1 sec., Br deriv., m. 201-2°; Br vapor gives no characteristic color

C. J. WEST

Organic arsenic compounds. I. J. KLIPPEL. *Rozniki Chem.* 10, 777-83 (1931) (English) (1930). Methylphenyl α -naphthylarsine, $C_{10}H_7As$, m. 60.0-0.5° was obtained by the action of $PhMgBr$ on $Ph(C_6H_5)_2AsCl$. On addn of Cl or CNBr to this compd., and after decomposition of the addn products by heating phenyl α -naphthylchloroarsine, $C_{10}H_7AsCl$, m. 46.0-6.5°, and phenyl α -naphthylcyanoarsine, $C_{10}H_7AsCN$, m. 55.5-100°, resp. were formed. The chloride and cyanide were converted by hydrolysis into phenyl- α -naphthylarsine oxide, m. 116.5-7.5° which on oxidation by H_2O_2 gave the arsenic acid, m. 189.0-9.5°. The chloride and oxide described by Blicke and Smith (*C. A.* 24, 83) as thick non-crystg. oils are cryst. when in a pure condition

J. KLČERA

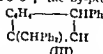
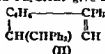
The action of bromine on naphthylamine- and aminonaphtholsulfonic acids. GUSTAV HELLER, HENRIK ARNOLD and JOHANNES SCHMIDT. *Z. anorg. Chem.* 43, 1132-7 (1930), cf. *C. A.* 22, 3653. 1,4-, 1,5- (I) and 1,8- $NH_2C_{10}H_6SO_3H$ react with an excess of Br_2 , yielding the corresponding 2,4-di-Br deriv., while the 2,5-, 2,6- and 2,7- compds. give the 1-Br deriv. The products from the 2,6- and 2,7- compds. were not characterized since the reaction in these cases was not under control. 2,3,6- (II) and 2,6,8- $NH_2C_{10}H_6SO_3H$ (III) yield dibromonaphthalenesulfonic acids of unknown structure. In each case I, SO₃H group has been substituted for a Br atom. In the product from II, a Br atom occupies position 1. Unknown mono-Br deriv. of I and III were also prepd. SO₃H deriv. of 1,2-, 1,8-, 2,5- and 2,8- $NH_2C_{10}H_6SO_3H$ yield bromonaphthoquinones upon reaction with Br_2 , the halogen replaces the SO₃H group when an excess is used. 2,5,7- (IV) and 2,6,8- $NH_2(OH)C_{10}H_6SO_3H$ and 2,8,3,6- $NH_2(OH)C_{10}H_6SO_3H$ aff. yield 1,3,6,7-tetrabromo-2-amino-5,8-naphthoquinone (V), m. 211°, Ac deriv. m. 255°, di-Ac deriv. m. 160-1°. V reacts with $PhNH_2$, yielding the anilino-tetrabromo deriv., $C_{10}H_6O_2N_2Br_4$, m. 215-6°. Upon reduction of V with $SnCl_2$ 1 of the Br atoms is lost, the product $C_{10}H_6O_2NBr_3$, m. > 300°. IV also reacts with 2 mols of Br_2 , the probable product being 1,3 dibromo-2,5,7 aminonaphtholsulfonic acid, 1,8,4- $NH_2(OH)C_{10}H_6SO_3H$ and 1,8,4,6- $NH_2(OH)C_{10}H_6SO_3H$ (VI) both give 2,4,7 tribromo-1-amino-5,8-naphthoquinone, while 2,4,6,7 tetrabromo-1-amino-5,8-naphthoquinone, m. 255° (Ac deriv. m. 184°) is derived from both 1,8,5- $NH_2(OH)C_{10}H_6SO_3H$ and VI. VI also yields 2,4,8-tribromo-1,8,6-aminonaphtholsulfonic acid and the K salt of 1,2,4 tribromo-5,8-naphthoquinone-6-sulfonic acid. This last product indicates that the NH_2 group has been replaced by Br. It was isolated with 2.5 mols of H_2O , 1.5 mols of which was

driven off by heating at 150°. The bromination of 1,8,3,6-NH₂(OH)C₁₂H₁₁(SO₃Na)₂ in KCl soln gives the K salt of 2, 4(2)-dibromo-1-amino-5,9-naphthoquinhydrone-3,6-disulfonic acid, the reduction of which compd leads to the formation of the hydroquinone compd, C₁₂H₁₁O₂NS₂Br₂. By heating 0.6 g of 1,2,4-NH₂(OH)C₁₂H₁₁SO₃H (VII) with 1 cc of 50% Br₂ in H₂O at 100°, 0.4 g of 3,4,1,2-Br₂C₁₂H₁₁(O)₂ was isolated. When a free flame was used in heating VII with H₂O, 3,2,1,4-Br₂(HO)C₁₂H₁₁(O)₂ was obtained. 1,2,4,6-NH₂(OH)C₁₂H₁₁(SO₃H)₂ yields the K salt of 4-bromo-1,2-naphthoquinone-6-sulfonic acid. I. M. LIVING

Reactivity of substituents in the C₉-nucleus. VINCENT JACOBI *J prakt Chem* 129, 55-96(1931).—Because cyclopentadiene is not only difficult to prep but polymerizes quickly at room temp, indene (I) has been used in this study. Much of the chemistry of I is reviewed. Details are given of the reduction of I and the prepn of its Li deriv. Distn of dibromohydrindene at ordinary temp gives 5-bromoindene b₁₁ 120°, m 36°, a trace of concd H₂SO₄ gives an intense bluish red color, oxidation gives 4-BrC₁₀H₇(CO₂H)₂. Diketohydrindene (II) and p-C₁₂H₄(ClHO)₂, heated at 110-20°, give p-aldehydobenzalindandione, C₁₂H₄(CO)₂C₆H₄(CHO)₂, pale yellow, m 173°, concd H₂SO₄ gives an intense orange color, PhNH₂ gives p-C₁₂H₄(Cl N-NHPh)₂. With twice the amt of II, there results terephthalidenebisdiketohydrindene, m 293°, a by-product, yellow m over 300°. II and 1 mol MeMgI give unchanged II indicating the structure C₁₂H₄ C(OH) CH CO, even 2 mols MeMgI give unchanged

II; 5 mols MeMgI give 1,7-dihydroxy-3-methylindene b₁₁ 120°, repeated vacuum distn or distn at atm pressure splits off 1 mol H₂O, giving α-methylindone. The 5-Br deriv. of I and Mg in Et₂O give the Grignard reagent after heating several days; with H₂O it gives I. 1,2,5-Tribromohydrindene (5-bromoindene dibromide), oil which decmps very easily, cooking with 10% EtOH gives 5-bromoindene hydroxyamide, m 80-5-15°. I dibromide, b₁₁ 143.5-4.5°, m 31.5-2.5°, gives with 10% EtOH 2-ethoxycarbonylhydrindene, oil; the 2-MeO deriv is also an oil. Distn of the EtO deriv in vacuo gives 1-bromoindene, b₁₁ 135.5-6°, concd H₂SO₄ gives an intense ruby-red color the Mg deriv with CO₂ gives indene-1-carboxylic acid, m about 70° (decompn). Distn of indene dibromide at atm pressure gives 3-bromoindene, b₁₁ 110-20°; Mg in Et₂O followed by CO₂ gives the 3-carboxylic acid, m 160° (decompn). A by-product is 3-indenophenone (dindyl ketone), m 233°. Indene is also formed in small amts. 1,2,3-Tribromohydrindene (3-bromoindene dibromide), oil which readily decmps, 10% EtOH gives 3-bromoindene hydroxybromide, m 90°. I and Br in H₂O, boiled 3 hrs, give a tribromoindene, pale yellow, m 133.5-4°. HNO₃ gives a dibromophthalic acid, showing that 2 nl the Br atoms must be in the benzene ring. C. J. WEST

Alkali organic compounds. OTTILIE BLUM-BERGMANN *Ann* 484, 26-51(1930).—Schlenk and Bergmann (*C. A.* 22, 4495), by the action of Na upon 1-benzohydrindene 3-phenylindene (I), obtained 2 dihydro derivs, m 174.5-5° (II) and 133-4.5° (III), Wieland and Kloss (*C. A.* 23, 3696), by the reduction of 1-benzohydrindene-3-phenyl 2-chloroindene (IV), obtained a 3rd isomer (V), m 115°. In order to clear up this matter, II and V have been synthesized. IV is dimorphous, crystg from benzine as orange-red cubes, m 147-8°, and red-violet needles, m 157-8°; reduction with H₂ and red P in AcOH gives V, needles, m 115°, or rhombic leaflets, m 130-1.5°; the leaflets are changed into the needles on repeated crystn. The Li deriv (VI) of 3-phenylindene and BzPh give I. VI and ClCO₂Me give di-Me 3-phenylindene-1,1-dicarboxylate, m 105-7°; catalytic reduction gives the dihydro deriv, m 102-3.5°; the free dicarboxylic acid m 175-7° (decompn), heated above its m p, it gives 3-phenylhydrindene-1-carboxylic acid, b₁₁ 227-9°, m 143-4.5°; Me ester, b₁₁ 201.5-5°, the latter with PhMgBr gives 3-phenyl-1-(α-hydroxybenzohydrindyl)hydrindene, m 122-4°, which loses H₂O when heated with AcCl, giving V. The Li deriv of fluorene and MeI give 9,9-dimethylfluorene and the 9-Me deriv. The Li deriv of indene and Ph₂CHBr give 1-benzohydrindene, m 163-4°. VI and Ph₂CHBr give II, m 174.5-5°; the by-products



include (Ph₂CH)₂ and 3,3'-diphenyl-1,1'-dindenyl; a dimer of 3-phenylindene, yellow, m 207-9°, and a compd C₂₂H₁₂, leaflets, m. 166-7°, and yellow prisms, m 162-4.5°; reduction of the last gives the compd C₂₂H₁₄, m 182-4°. The Li deriv of 2-phenylindene and Ph₂CHBr give 1-benzohydrindyl-2-phenylindene, m 175°, whose dihydro deriv m 151°. This shows that the compd C₂₂H₁₂ above is not produced by the wandering

of the Ph group. The action of Al Hg upon I gives a mixt of II and III. Reduction of III with Na and AmOH gives I benzohydryl-3-phenylhydrazide, m 133-5°, also obtained by reducing V. In the reduction of anthracic acid, there results a double compd of this acid with the tetrahydro acid, $C_{24}H_{18}O_2$, m 190-2°, green luster. The action of CO_2 upon the Li deriv of anthracene gives 9,10-dihydro 9,10,10-tricarboxylic acid, crystg with 1 mol H_2O , m 290-92° (decompn), heating 13 hrs at 100° splits off the H_2O , further heating for 14 days also splits off 1 mol CO_2 , *tri-Me ester*, m 147.5-8°. Heating the acid in AcOH gives α , 9,10 dihydroanthracene 9,10-dicarboxylic acid, m 162-3.5°. Ph_3CNa and $CO(OPh)_2$ give Ph_3CCO_2Ph , m 124-7°, $CO(OEt)_2$ gives Ph_3CCO_2Et , m 118-20°, $(Ph_3CNa)_2$ and $CO(OPh)_2$ give $(PhC)_4$. The Li deriv of fluorene and $CO(OMe)_2$ give Me biphenyleneacetate and dibiphenyleneacetone, m 230-2° (decompn). Ph_3CHNa and $CO(OMe)_2$ give $(Ph_3CH)_2CO$, m 134-45°. $(Ph_3CNaClf)_2$ and $CO(OMe)_2$ give 1,1,4,4-tetraphenylcyclopentan-5-one. The di Li deriv. of anthracene and $CO(OMe)_2$ give bisdihydroanthracyl ketone, m 238-40° (evolution of CO); in its thermal decompn, 88.5% of the theory of CO_2 is evolved and a mixt of anthracene and its dihydro deriv is formed. The Na deriv of anthracene behaves in the same way $(Ph_3CH)_2CO$ with Na in Li_2O gives the Na salt of the enolate, $C_{24}H_{18}ONa$, dry distn gives tetraphenylallene and 1,1,3,3-tetraphenylpropylene. C J WEST

Coupling of diazonium compounds with 2-methoxy-3-naphthoic acid G B JAMBULSERWALA AND F A MASON *J Soc Dyers Colourists* 46, 339-41(1930).—The Me ether of β hydroxynaphthoic acid will couple in H_2O with diazotized *p*-, *o*-, and *m* nitroaniline, 4- and 5-nitro-*o*-anisidine, sulfanilic acid, benzidine and tolidine. MILTON HARRIS

Keto-oxidation product of α -acetylacenaphthene. K. DZIEWONSKI AND J REISS. *Bull. intern. acad. polonaise* 1930A, 62-5.—Contrary to Graebe and Haas (*Ann* 327, 77-103(1903)) oxidation of 5-acetylacenaphthene (best prepd by the action of $AlCl_3$ on $AcCl$ and acenaphthene without solvent) with $Na_2Cr_2O_7$ and AcOH at 50° gives only traces of 4,1,8- $C_{12}H_7Ac(CO_2H)_2$, the main products being 5-acetylacenaphthenequinone, m 192-3° [bisphenylhydrazone, m 240-2°, *dioxime*, m 275-8° (decompn)], and 2,2'-diketo 5,5'-diacetyldiacenaphthylidene, m 205° (decompn). B C A

Derivatives of 3-bromoacenaphthene. K DZIEWONSKI, MILE J SCHOEN AND MILE A GLAZNER *Bull intern acad polonaise* 1929A, 636-49.—Oxidation of 3-bromoacenaphthene by $Na_2Cr_2O_7$ in AcOH at 30° yields, beside a 4 bromonaphthalic acid 3-bromoacenaphthenequinone, m 235-6° (Graebe, *Ann* 327, 77-103(1903), gives m p 194°) [monophenylhydrazone, m 179-80°, bisphenylhydrazone, m 225-6° (literature 153° and 134°, resp)], *dioxime*, m 230-1° (decompn), 3 bromoacenaphthaphenazine, from the foregoing quinone and o $C_{12}H_7(NH)_2$, m 261-3°, and 3,3'-dibromobiacenedione, $(C_{12}H_7Br)_2 \begin{smallmatrix} C \\ \diagup \quad \diagdown \\ CO \end{smallmatrix}$, m 320-1° (for nomenclature, cf C A 20, 1234).

Nitration of 3 bromoacenaphthene in AcOH gives mainly 3 bromo-4 nitroacenaphthene, m 159-61°, converted by $Na_2S_2O_4$ into 3 bromo-4 aminoacenaphthene, m 133°. The orientation of these derivs is established by the further reduction of the latter, by means of Na Hg and alc, to the known 3 aminoacenaphthene. 3 Bromo-4 nitroacenaphthene is oxidized by $Na_2Cr_2O_7$ to 4 bromo 5 nitronaphthalic acid, m 295° (anhydride, m 312°, *Me ester*, m 162°). Two isomeric SO_3H acids, termed α and β resp, are obtained by the interaction of $ClSO_3H$ and 3-bromoacenaphthene at the ordinary temp, and are sep'd by means of their Na salts. 3-Bromoacenaphthene α -sulfonic acid forms the more sparingly sol Na salt (aniline salt, m 260-1°, β naphthylamine salt, m 263-6°, chloride, m 134-5°, amide, m 137-8°), oxidation affords 3 bromo α sulfonaphthalic anhydride, isolated as its Na salt, 3 Bromoacenaphthene β sulfonic acid (aniline salt, m 256-7°, chloride, m 192-3° amide, m 233-4°, *Et ester*, m 140-1°) yields a similar naphthalic acid on oxidation. Sulfonation with H_2SO_4 (d 1.84) at 80-90° converts 3 bromoacenaphthene into a disulfonic acid isolated as the Ba salt the di Na salt (+3 H_2O) is readily sol in water [dichloride m 181-2°, diamide, m 289° (decompn), *Et ester*, m 164°]. B C A

Reduction products of the hydroxyanthraquinones. XII. GORDON F. ATTREE AND ARTHUR G PERKIN *J Chem Soc* 1931, 144-73, cf C A 24, 2456.—2 Methoxyanthraquinone (I) in H_2SO_4 and Cu powder, gradually added and heated to 140° for 1 hr, and the resulting product treated with Ac_2O and C_6H_5N , give the acetate, m 152-4°, of 3 methoxy-9-anthranol (II), m 108-9°. II with $FeCl_3$ -AcOH and EtOH gives about 30% of 3,3'-dimethoxydianthrone, m 215-7° to a green liquid, concd H_2SO_4 gives a pale yellow soln, the main product of the reaction is I. 3,3'-Dimethoxydianthranyl 9,9'-diacetate, m 228-30°, sol in H_2SO_4 with a violet red tint, I in C_6H_5N

transforms this to 3,3'-dimethoxyanthraquinone, yellow, m 241-5°, the latter in $\text{C}_{12}\text{H}_5\text{N}$, exposed to light for 2 weeks, gives 2,2'-dimethoxy-naphthanthrone, orange-red, does not m 310° and gives a red soln in H_2SO_4 . Isomethoxyanthrone (III), reduced by $\text{SnCl}_2\text{-HCl}$, gives 3,6-dihydroxy-9 (isomethoxyanthrone)anthrol in 183-1° (Roemer and Schwieger, *Ber.* 15, 1011 (1882), give 153°), FeCl_3 in AcOH and LiOH gives 3,3',6,6'-tetraacetoxy-9,9'-dianthranil diacetate, pale yellow, m 280-2°, hydrolysis with HCl-AcOH gives 3,3',6,6'-tetrahydroxydianthrone, pale yellow, darkens 250°, turns green 300°, and does not lose a definite m p. The hexameteite does not give a dianthronquinone with 1 m $\text{C}_{12}\text{H}_5\text{N}$. III gives a di-Me ether (IV), yellow, m 215-7°, with Cn and H_2SO_4 , this yields the acetate, pale yellow, m 180-1° of 3,6-dimethoxy-9-anthranol, m 158-60°, the latter with FeCl_3 in AcOH gives 3,3',6,6'-tetraacetoxy-9,9'-dianthrone, m 242-3°, whose diacetate, yellow, m 255-6°, does not react with 1 m $\text{C}_{12}\text{H}_5\text{N}$. IV and $\text{SnCl}_2\text{-HCl}$ in AcOH give 3-methoxy-6-acetoxy-9-anthranil acetate, m 107-9°, sol in H_2SO_4 with a bluish red color, becoming yellow, hydrolysis with HCl-AcOH gives 6-hydroxy-3-methoxyanthranol, m 211-6°, giving a yellow soln with a green fluorescence in H_2SO_4 , oxidation of the acetate with CrO_3 in AcOH gives the mono-Me ether of III, yellow, m 283-5° as the *1c* deriv., m 194-6°, giving a crimson color in H_2SO_4 , the mono-Me ether also results in very small yields in the action of Me_2CO on III. 6,6'-Dihydroxy-1,1'-dimethoxyanthrone in 202-5° 3,3'-Dihydroxy-9-anthranil and $\text{FeCl}_3\text{-AcOH}$ give 3,3',6,6'-tetrahydroxydianthrone in 415-8° sol in H_2SO_4 with an orange yellow color, the tetraacetoxy diacetate darkens about 241° but does not have a definite m p. H_2SO_4 gives a scarlet color changing through brown to green. 2,2',3,3'-Tetrahydroxydianthrone, from hysteronin (V) through acetyl-hysteronanthrol, with Ac_2O gives the *tetra-Ac* deriv., pale yellow, m 251-7° with $\text{C}_{12}\text{H}_5\text{N}$ there results the *tetra-Ac* diacetate, m 241-5°, sol in H_2SO_4 with a red color, changing to green and then brown. The di-Me ether of V, with $\text{SnCl}_2\text{-HCl-AcOH}$, gives the acetate, m 161-71°, of 2,1-dimethoxyanthranol, pale yellow, m 111-6°, with $\text{FeCl}_3\text{-AcOH}$ there results 2,2',3,3'-tetramethoxyanthranol, m 213-5°, sol in H_2SO_4 with an orange red color, the diacetate m 104-6°. Oxidation of the anthrone with $(\text{NH}_4)_2\text{S}_2\text{O}_8$ gives 2,2',3,3'-tetramethoxydianthraquinone (VI), pale yellow, m 280-1°, sol in H_2SO_4 with a dull violet color, quickly passing to a blackish brown. C_{12}H_5 gives a deep green soln. Exposure to light in $\text{C}_{12}\text{H}_5\text{N}$, LiOH for 20 days gives 2,2',3,3'-tetramethoxyhelianthrone, m 287-8°, sol in H_2SO_4 with a deep unchangeable color, HCl-AcOH at 100° for 3 hrs. splits off the MeO groups, giving the *tetra-HO* deriv., iridescent metallic leaflets; *tetra-Ac* deriv., orange, m 215-7°, CrO_3 and AcOH oxidize this to 2,2',3,3'-tetraacetoxy-1,1'-dianthraquinonyl, pale yellow, m 200°, and then 268-70°, H_2SO_4 gives a red brown soln. Hydrolysis gives the *tetra-HO* deriv., dull yellow, does not m 360° and gives a green blue soln in alkalis and a brick red color in H_2SO_4 , it dyes Sn , Al , Fe and Cr -mordanted calico, producing bright orange, dull red purple, purple and dull brown shades. VI in $\text{C}_{12}\text{H}_5\text{N}$ and LiOH , exposed to the light for 6 weeks, gives 2,2',3,3'-tetramethoxynaphthanthrone, orange needles; the *tetra-HO* deriv. forms scarlet needles, sol in H_2SO_4 with a violet red color, it possesses little affinity for mordanted calico and is a poor dye toward wool mordanted with Sn , Al , Fe or Cr , *tetra-Ac* deriv., yellow, does not m 360°. 4,4'-Dihydroxydianthrone, m 256-8°, sol in H_2SO_4 with an orange red color; the diacetoxy diacetate m 205-7°. 1-Methoxyanthrone m 329-331°, $\text{FeCl}_3\text{-AcOH}$ gives 1,1'-dimethoxydianthrone, pale yellow, does not m 330°; the diacetate, pale yellow, m 290-1°, the crimson soln in H_2SO_4 changes to violet and then blue on standing. Iodine and $\text{C}_{12}\text{H}_5\text{N}$ give 1,1'-dimethoxydianthraquinone, pale yellow, m 315-6°; oxidation of a colloidal soln with alk. $\text{K}_2\text{Cr}_2\text{O}_7$ did not have any effect; however, exposure of a Me_2CO soln to the light for 1 month gives 4,4'-dimethoxyaphthalanthrone, orange red, does not m 360° and gives with H_2SO_4 a highly fluorescent, crimson soln; HCl and AcOH give the corresponding 4,4'-di-HO deriv., scarlet needles, does not m 360°. The other product of the action of light is 4,4'-dimethoxyhelianthrone, salmon, does not m 360°, gives a violet color in H_2SO_4 , the 1,1'-di-HO deriv., orange red needles, was fused in alkalis and does not appear to give an *Ac* deriv. Anthracene di-Me ether (VII) and $\text{SnCl}_2\text{-HCl}$ in AcOH give 1,5-dimethoxyanthrone, brown-red, m 181-2°, giving a red soln in H_2SO_4 , quickly changing to green blue; acetate, pale yellow, m 169-71°. The anthrone and FeCl_3 in AcOH give 1,1',5,5'-tetramethoxydianthrone, m 305-7°, which is not attacked by Ac_2O and $\text{C}_{12}\text{H}_5\text{N}$ or by 1 or $(\text{NH}_4)_2\text{S}_2\text{O}_8$. VII and $\text{SnCl}_2\text{-HCl}$ at 65° for 3 hrs. give 10-hydroxy-1,5-dimethoxyanthrone, violet, m 151-5°, sol in H_2SO_4 with a pure blue color; boiling the mixt. for 1 hr. gives 1-hydroxy-5-methoxyanthrone, pale yellow, m 121-3°, 1-acetoxy acetate, yellow, m 161-3°, the LiOH soln has a strong blue fluorescence. 1,1'-Dihydroxy-5,5'-dimethoxyanthrone, pale yellow, m 287-9°, sol in H_2SO_4 with an

orange-yellow color, quickly passing to greenish blue, 1,1'-diacetoxy diacetate, pale yellow, m 250-2°. Chryzarin di-Me ether (VIII) yields 1,8-dimethoxyanthrone, yellow, m 196-7°; air oxidation of an alk. soln gives 1,1',8,8'-tetramethoxydianthrone, does not m 340°, gives a red brown color in H₂SO₄, rapidly changing to dull green, is not changed by prolonged digestion with Ac₂O and C₁₂H₅N. SnCl₄-HCl and VIII give 1-hydroxy-8-methoxyanthrone, pale yellow, m 183-5°, the 1-acetoxy acetate, m 164-5°. 1,1'-Dihydroxy-8,8'-dimethoxydianthrone, pale yellow, does not m below 330°, Ac compd, m 244-7°. Alizarin 2-Me ether and SnCl₄-HCl give 1-hydroxy-2-methoxyanthrone (IX), oxidized to 1,1'-dihydroxy-2,2'-dimethoxydianthrone, yellow, m 195-7°, the 1,1'-diacetoxy diacetate m 242-5°; IX and I in C₁₂H₅N give 4,4'-dihydroxy-3,3'-dimethoxydianthrone, iridescent magenta like leaflets, does not m 330°, sol in H₂SO₄ with an intense greenish blue color, various other methods of prepn are given. HCl and AcOH give the *tetra HO* deriv, maroon-colored, does not m below 300°, H₂SO₄ gives a deep blue color, NH₄OH gives a violet color, it does not dye mordanted calico but has weak tinctorial power toward mordanted wool, Ac₂O and C₁₂H₅N give 4,4'-dihydroxy-3,3'-diacetoxydianthrone, iridescent, chocolate-colored needles, m about 325°. 1,1',2,2'-Tetrahydroxydianthrone, pale yellow, m 240-1°, *tetra Ac* deriv, m 245-7°. Purpurixanthin di-Me ether, on reduction, gives 10 hydroxy 1,8-dimethoxyanthranol, red, m 156-8°, reduction of the di-Me ether with SnCl₄-HCl in hot AcOH gives a *tri Ac* deriv, yellow, m 154-5°, of 1,3-dihydroxy (purpurixanthinanthrone, m 217-9°. 3,4-Dimethoxy-9-anthranol acetate, yellow, m 146-8°; the *Me* ether yellow, m 116-8°.

C. J. WEST

The action of organomagnesium compounds upon furanacetophenone. New ketones in the furan ring. N. N. MAXIM. *Bul. soc. chim. Roumanie* 12, 24-7 (1930), cf C A 25, 513. —M allowed organomagnesium compds to react upon furanacetophenone and always obtained catd ketones in the furan ring according to the following equation: $2 C_6H_5OCH_2CHCOPh + RMgX \rightarrow 2 C_6H_5OCH_2CHCOPh$. The furanacetophenone was prepd according to Kostanecki and Podjanski (*Ber* 29, 2215) by allowing furfural to act upon PhCOMe in the presence of dil NaOH. From this, M prepd β -furylbutyropheneone, MeCH(C₆H₅O)CH₂COPh, β -furylpropiophenone, PhCH(C₆H₅O)CH₂COPh and β -phenyl β -furylpropiophenone. These ketones have been characterized by their semicarbazones and oximes. The yield is almost quant. This reaction can be used as a general method of prep catd ketones in the furan ring. The mechanism of the reaction will be published elsewhere.

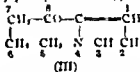
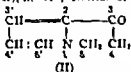
F. R. G.

Preliminary work to ring syntheses of porphyrins, etc. IV. A bromomethylpyrrole and its reactions. HANS FISCHER AND Oskar SOHN. *Ann* 484, 113-30 (1930), cf C A 25, 961. —2,4-Dimethyl-5-carbethoxypyrrole-3-acetyl acid (I) and Br in CS₂ give a dibromide, C₁₁H₁₁NOBr₂, which does not have a sharp m p, the Br is labile and with Me₂CO on the H₂O bath 60% I is regenerated, with EtOH-H₂O on the boiling H₂O bath there results 2,4-dimethyl-5-carbethoxy-3- ω -bromomethylpyrrole (II) m 155° (decompn). Heated with HI-AcOH at 100° for 2 hrs, II is transformed to dimethylpyrrole, catalytic reduction of II gives cryptocarbethoxypyrrole. Br gives a dibromide AgCN in MeOH gives 2,4-dimethyl-5-carbethoxy α - β -3-dimethoxyethylpyrrole m 112°. Ag₂O Ag or Cu gives the same product, with EtOH in place of MeOH there results 2,4-dimethyl-5-carbethoxy-3- ω -vinyl ether, m 87°. One mol II and 1 mol SO₂Cl₂ give a mono-Cl deriv, C₁₁H₁₁O₂NCIBr, m 168°. 2 mols SO₂Cl₂ give 2 formyl 4 methyl-5-carbethoxy-3-bromomethylpyrrole (III) m 140°, the aldehyde m 230° (decompn), the phenylhydrazine m 142° (decompn). Condensation of III and cryptopyrrole with HBr give the methene HBr, C₁₁H₁₁O₂N₂Br₂, brick red blackens 170° but does not m 300°. Hydrolysis of III gives 2 formyl 1-bromomethyl 4 methyl-5-pyrrolecarboxylic acid m 238° (decompn) which condenses with cryptopyrrole to a methene HBr, C₁₁H₁₁O₂N₂Br₂, reddish brown needles. Three mols SO₂Cl₂ and II give a di-CO₂H acid, whose a *Me* α' *Et* ester C₁₁H₁₁O₄N₂Br m 131°, di *Et* ester m 115°, the free acid C₁₁H₁₁O₄N₂Br slowly carbonizes above 230°. Catalytic reduction gives 3 ethyl-4 methylpyrrole-2,5-dicarboxylic acid. The *Et* ester (IV) of I, m 134° results from I, EtOH and HCl or by condensation of 2,4-dimethyl-5-carbethoxypyrrole and H₂O-CCl₃-CO₂H with PhNH₂, the dibromide m 121° (decompn), the Br is removed by Zn in AcOH, giving IV, while catalytic reduction gives *Et* 2,4-dimethyl-5-carbethoxy-3-propionate m 73°. EtOH KOH gives 2,4-dimethyl-5-carbethoxy-3- ω -carbethoxy- ω -ethoxyvinyl pyrrole, m 136° and a Br-contg acid m 196° (decompn), with MeOH a compd, C₁₁H₁₁O₄N₂, m 178° results. Oxidation of I with CrO₃ gives 2,4-dimethyl-5-carbethoxypyrrole-3-carboxylic acid m 271°.

C. J. WEST

New derivatives of pyrrole. I. Synthesis of 3-keto-4,5-dihydro-di-(1,2)-pyrrole and of 8-keto-5,6,7,8-tetrahydropyrrocoline. GEORGE R. CLEW and GEORGE R.

RAMAGN *J. Chem. Soc.* 1931, 49-53.—K pyrrole and $\text{ClCH}_2\text{CO}_2\text{Et}$ in C_6H_6 heated 0.5 hr at 100° , give the *Et* ester, b_p 112° , of 1-pyrrolylactic acid, m 91° ; amide, m 169° . β -1-Pyrrolylpropionic acid, m 62° . *Et* ester, b_p 122° , amide, m 81° . Attempts to effect ring closure were unsuccessful. It was also not possible to effect ring closure with succinobutylamide, b_p 110° . K pyrrole and $p\text{-MeC}_6\text{H}_4\text{SO}_2\text{CH}_2\text{CH}_2\text{Cl}$ give 1- β -chloroethylpyrrole, b_p 84° , β -cyano deriv (I), b_p 140° , 1- γ -chloropropylpyrrole, b_p 87° . The last compd is recovered unchanged after prolonged boiling with EtOH-KOH . AlCl_3 also failed to effect ring closure. Passing dry HCl through I in dry 1:1 EtO contg ZnCl_2 gives 3-keto 1,5-dihydro- δ -(1,2) pyrrole (II), m 61° , best isolated through the semicarbazone, m 211° , piperonylidene compd, pale yellow, m 191° . $\text{I O}_2\text{CCCH}_2\text{CH}_2\text{COCl}$ and K pyrrole give 1,1- β -1-pyrrolylpropionate, m 50° , treatment with Na in $\text{C}_6\text{H}_5\text{Me}$, and distn with CaCl_2 failed to remove 1 OEt . Levulyl chloride and K pyrrole give β -1-pyrroylethyl Me ketone, b_p 144° (phenylhydrazone, m 131° , semicarbazone, m 190°), treatment with acids and alkalis always resulted in amide hydrolysis. γ -1-Pyrrolylbutyronitrile, b_p 152° , with ZnCl_2 and HCl , gives 8-keto-5,6,7,8-tetrahydro-pyrrocoline (III), m 31° , semicarbazone, m 193° , piperonylidene deriv, yellow, m 136° .



C. J. WIST

Cyclotellurobutane (tetrahydrotellurophen). GILBERT T. MORGAN AND FRANCIS H. BURSTALL. *J. Chem. Soc.* 1931, 180-4.— $(\text{CH}_3)_2\text{Te}$ and amorphous Te, heated at $130\text{--}40^\circ$ for 5 hrs, give cyclotelluributane 1,1-diodide (I), purple plates or bright red prisms, m $140\text{--}50^\circ$; reduction with SO_2 in boiling H_2O gives cyclotellurobutane (II), b_p $105\text{--}6^\circ$, b_m $160\text{--}7^\circ$, n_D^{20} 1.6175, penetrating and unpleasant odor, oxidizes readily in the air, reacts explosively with concd HNO_3 and gives a red color with H_2SO_4 ; the mercurichloride m $146\text{--}7^\circ$. II and Cl give cyclotelluributane 1,1-dichloride, m $111\text{--}2^\circ$, dibromide (III), m 130° , decomps at 210° . III and IV in equimol proportions in H_2O or II and Na_2CO_3 or NaOH sufficient to remove half the Hr , give the 1,1'-oxydibromide, $[(\text{CH}_3)_2\text{TeBr}]_2\text{O}$, m 207° (decomps); alkalis convert this into the 1-oxide IV, $(\text{CH}_3)_2\text{TeO}$, m 211° (decomps), also obtained by the oxidation of II or the action of excess alkali on the dihalides; halogen acids regenerate the corresponding dihalides. II and MeI give cyclotelluributane 1-methiodide, vaporizes without melting at 210° . $(\text{CH}_3)_2\text{Br}$ and Al telluride, heated at 125° for 2 hrs, give 1,1-bromobutylcyclotelluributane 1-bromide, $(\text{CH}_3)_2\text{Te}[(\text{CH}_3)_2\text{Br}]_2$, obtained from the Me_2CO and 1 EtOH exts of the reaction product, m $152\text{--}3^\circ$; above 155° it dissociates into $(\text{CH}_3)_2\text{Br}$ and II; the corresponding iodide m $175\text{--}6^\circ$ (decomps), the aq ext of the above reaction product contains tetramethylene- α,δ -bis-cyclotelluributane 1,1'-dibromide, $(\text{CH}_3)_2\text{TeBr}[(\text{CH}_3)_2\text{TeBr}(\text{CH}_2)_4]$, crystg with 1 mol H_2O , m 225° with dissociation to $(\text{CH}_3)_2\text{Br}$ and II; diiodide, deep yellow, decomps 215° .

C. J. WIST

Heterocyclic systems containing selenium. IV. Cycloselenohexane. GILBERT T. MORGAN AND FRANCIS H. BURSTALL. *J. Chem. Soc.* 1931, 173-80, cl. C. J. 24, 4237.— $(\text{CH}_3)_2\text{Br}$ and Na_2Se , refluxed 2 hrs, give a mixt. of yellow polymers and a very small quantity of cycloselenohexane (I), b_p $105\text{--}6^\circ$, b_m $188\text{--}90^\circ$, n_D^{20} 1.6470, d_4^{20} 1.353, d_4^{25} 1.378—0.001011; the concd H_2SO_4 soln develops yellow to red tints on warming, I is not oxidized in the air; mercurichloride, m $193\text{--}4^\circ$ (decomps). I and 1 in 1:1 EtOH give cycloselenohexane 1,1-diodide, reddish purple, m 82° , 1,1-dibromide, yellow, m $118\text{--}9^\circ$; it slowly decomps giving $(\text{CH}_3)_2\text{Br}$ and Se, 1,1-dichloride, m 78° ; 1-methiodide, m $149\text{--}50^\circ$ (decomps). The more sol fraction of the polymerides (in petr. ether, b_p $60\text{--}80^\circ$) is a dimer, $(\text{C}_6\text{H}_{12}\text{Se})_2$, m 92° , depolymerizes above 220° ; dibromide, pale yellow, m 145° (decomps). The less sol. polymer, $(\text{C}_6\text{H}_{12}\text{Se})_n$, is a waxy microcryst powder, m $30\text{--}7^\circ$; mol-wt detns. indicate approx 12 mol H_2 ; the bromide decomps slowly and m $91\text{--}5^\circ$ (decomps). Heating this polymer at 220° gives 2-methylcycloselenopentane (II), b_p 90° , b_m $160\text{--}71^\circ$, n_D^{20} 1.5205, d_4^{25} 1.287, d_4^{25} 1.3115—0.001111; the dichloride is a colorless liquid, the dibromide a reddish yellow gum and the diiodide a deep red gum; mercurichloride, m 112° ; 1-methiodide, m 104° (decomps). The constitution of II is shown by heating the dibromide, which gives $\text{Br}(\text{CH}_2)_4\text{CHMeBr}$, which was converted into the dinitrile and α -methylpimelic acid $(\text{CH}_3)_2\text{Br}$, and KSeCN , refluxed in Me_2CO for several hrs, give hexamethylene α,δ -diselenocyanate, thick yellow oil which decomps on distn; oxidation by the air gives cyclodiseleno-

hexane, brown, highly polymerized, m about 40° and decomps at 250°, giving II; oxidation of either of these compds with HNO₃ gives hexamethylenediselenous acid, decomps 147°. The mean parachor of I (302.1) and of II (304.2) indicates that they are normal liquids.

C. J. WEST

Behavior of certain thiophenes in heptane and naphtha solutions. R. W. BOSE AND M. W. CONY *Ind Eng Chem* 23, 93-5 (1931).—Tetramethylene and pentamethylene sulfides have been studied along with Li₂S and thiophene, in heptane and three naphtha solns. HgCl₂, KMnO₄, H₂O₂, and MeI react with the thiophenes to form definite products easily identified and purified. Br and HgI₂ form unstable products. B and W state in general thiophenes resemble alkyl sulfides more than thiophene.

M. C. ROGERS

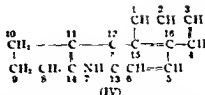
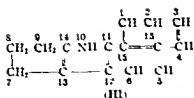
Indophenol. GUSTAV HELLER. *Chem-Ztg* 54, 985-7 (1930), cf. C. A. 19, 1277.—A theoretical discussion of the formula proposed by Steinkopf and Roch (C. A. 25, 292) for the constitution of indophenol.

JAMES WHITE

Syntheses of glucosides. VII. Synthesis of 6-bromomondicane. ALEXANDER ROBERTSON AND ROY B. WATERS. *J Chem Soc* 1931, 72-6, cf. C. A. 25, 1232.—Me 4-bromomondicane, m 78°, does not react with ClCH₂CO₂Me when heated alone or with AcONa or CaH₂N. 4-Bromomondic acid and ClCH₂CO₂H, heated in 5-5% aq KOH for several hrs and the resulting product esterified with MeOH and H₂SO₄ give di-Me 5-bromophenylglycine 2-carboxylate, m 101° (almost quant yield), Na in CaH₂ contg a little MeONa gives Me 6-bromo-3-hydroxyindole 2-carboxylate (I), m 192°, di-Ac deriv, m 151°. Refluxing the ester with MeOH NaOH in N for 15 hrs gives 6-bromomondic acid, sepg from H₂O as the trihydrate, m 198° (decompn), AcO and AcONa give 6-bromo-1-acetyl-3-acetoxyindole, crystg from dil EtOH as the hydrate, sinters 122°, m 150-2°. Adding aq KOH to I and O-tetraacetyl- α -glucosidyl bromide in Me₂CO at 10° gives Me 6-bromo-3-O-tetraacetyl- β -glucosidoryndole 2-carboxylate m 171°, [n]_D²⁰ -59.7° (Me₂CO), a penta Ac deriv could not be obtained. Hydrolysis by heating with MeOH KOH on the H₂O bath for 1 hr gives the K salt (II) of 6-bromo-3- β -glucosidoryndole-2-carboxylate, which forms a gel, prolonged hydrolysis by dil HCl gives 6,6'-dibromomondigotin and glucose. II, AcONa and Ac₂O, heated on the steam bath for 1 hr and then at 100° for 1 hr, give pentaacetyl 6-bromomondicane (5-bromo-1-acetyl-3-O-tetraacetyl- β -glucosidoryndole), m. 159°, [n]_D²⁰ -48.8° (Me₂CO), with dry NiH₂ in MeOH at 0° there results 6-bromo-3- β -glucosidoryndole (6-bromomondicane), crystg with 4 H₂O, m 64°, [n]_D²⁰ -6.4° (Me₂CO), drying over P₂O₅ for 60 hrs gives the semi hydrate, m 177°, the anhyd compd results by heating at 110° for 15 hrs and at 160° for 5 min. This is rapidly hydrolyzed by warm 2% HCl and by emulsion at 35-7°, with the liberation of glucose and 6-bromomondoxyl, partly oxidized by the air to 6,6'-dibromomondigotin. Hydrolysis by warm 3% HCl in the presence of isatin gives quant 6-bromomondigotin, purple with green metallic luster, does not m 240°, with p-O₂NC₆H₄CHO, there results p-nitrobenzaldehyde-6-bromomondigotin brick red, m. 237-8°.

C. J. WEST

Action of nitric acid on polycyclic indole derivatives. IX. STEPHEN A. BRYANT AND SYDNEY G. P. PLANT. *J Chem Soc* 1931, 93-105, cf. C. A. 24, 114, 1375.—In view of the widely different nature of the products obtained when HNO₃ acts on the acyl derivs of tetrahydrocarbazole, dihydropentindole, 7,8,9,10-tetrahydro- α,β -naphthacarbazole and 8,9,10,11-tetrahydro- α,β' -naphthacarbazole, and of the deep-seated variations observed in the reactions of substances formed in such cases by the addn. of HO and NO₂ or OH and OH to the double linkage, it is of interest to extend the investigation to certain closely related types. Cyclopentanone α -naphthylhydrazone (I), m 93°, the β -deriv (II), m 77°. I and dil H₂SO₄ heated 15 min at 100°, give 7,8-dihydro- α,β -naphthapentindole (III), m 167°, this is also formed by boiling I in glacial AcOH, it is picrate, reddish brown, m 167° (decompn) III and Ac₂O contg. a little H₂SO₄ boiled 10 hrs give the 10-Ac deriv, m 157°, and the 1(?) -Ac deriv, yellow, m. 215°, stable toward dil EtOH-KOH, oxime, m 236° (decompn) II and dil H₂SO₄ give 9,10-dihydro- α,β' -naphthapentindole (IV), m 103°, picrate, brown, m 189° (decompn) AcCl and IV in Me₂CO and aq NaOH give the 7-Ac deriv, m 170° with Ac₂O and H₂SO₄ this gives the 5(?) -7,4-Ac deriv, pale green, m 234°, also obtained from IV and Ac₂O, partial hydrolysis gives the 5(?) -Ac deriv, greenish yellow m 249°, yielding with BaCl₂ and KOH the 5(?) -acetyl-7 benzoyl deriv, pale greenish yellow, m 163°, which may again be partially hydrolyzed to the 5(?) Ac deriv, 7-Bz deriv, pale yellow, m 199°, 7-carbethoxy deriv, m 160°. The 7-Ac deriv, nitrated in AcOH, yields a 5(?) NO₂ deriv, lemon-yellow, m 217°, K₂CO₃ splits off the Ac group, giving the 5(?) -NO₂ deriv of IV, brick red, m. 228°.



The 7-Bz deriv. also yields a 5(1) NO_2 deriv. yellow, m 250° . The 7-carbethoxy deriv. yields a 5(1) NO_2 deriv. lemon yellow, m 202° , and also a 5(1) NO_2 deriv. yellow, m 220° (decompn.). 2-Methylcyclohexanone and $3\text{C}_6\text{H}_5\text{NHNH}_2$, heated 0.5 hr. and the resulting only hydrazone treated with dil. H_2SO_4 , give a mixt. of 1-methyl-5,9,10,11-tetrahydro- α' , β' -naphthacarbazole (V), m 115° (precipitate, dark crimson, m 201° (decompn.)), and 12-methyl-5,9,10,11-tetrahydro- α' , β' -naphthacarbazole, m 92° (picrate, bright yellow, m 110°). A mixt. of V, S and quinoline, boiled for 45 min., gives 1-methyl- α' , β' -naphthacarbazole, m 144° ; this was also synthesized from 2,3-HQC, $\text{H}_2\text{C}_6\text{H}_4$ and $\text{C}_6\text{Me}_5\text{NHNH}_2$ by heating at 150° and then distg. the 8-methyl- α' , β' -naphthacarbazole-4-carboxylic acid with soda lime. 5,6-Dihydro- α' , β' -naphthacarbazole (Ghigi, C. 24, 3797) with Ag_2O contg. a little H_2SO_4 , gives the 1(1) Ac deriv., yellow, m 233° (crime m 272° (decompn.)). Benzopentadiene gives a 7-Bz deriv., pale yellow, m 187° , with HNO_3 in AcOH there results 14-nitro-6-hydroxy-7-benzyl-6,14-dihydrobenzopentadiene, m 109° (decompn.). 20% KOH gives 14-nitro-6-hydroxy-6,14-dihydrobenzopentadiene, m 200° (decompn.). C. J. WEST

Diazotization in the pyrazole series. J. REILLY AND D. MACDOWNEY, *Proc. Roy. Irish Acad.* 39B, 497-504 (1930).—The diazonium salts of 4-aminopyrazole derivatives had been shown (C. 19, 1173) to be more stable than those of 5-aminopyrazoles contg. a Ph substituent. The stability of the diazonium chloride of 1-phenyl-3,5-dimethyl-4-aminopyrazole (I) showed this stability to be a function of the position of the NH_2 group and not to be due to the presence of the substituent Ph group. From the study of pyrazole/diazonium derivs., it appears that the essential factors for diazotization are

the grouping >CNH_2 and other loci of unsatn. I was prepd. by condensing (contro-

succetylacetone with PhNHNH_2 , HCl (Ann 325, 120-91) and reducing the resulting nitrosopyrazole with Sn and HCl . I forms a yellow, dihauc chlorophosphate, an unstable red ppt. with AuCl_3 and a Bz deriv., m $153-4^\circ$. It also condenses with BzH to give a benzylidene deriv., m. $90-1^\circ$. 4-Amino-3,5-dimethylpyrazole gives characteristic colors with phenols in alk. soln. and its diazonium chloride forms a triazo deriv. with NaN_3 , but I fails to give these reactions. Treatment of an acid soln. of I in dil. EtOH with EtNO_2 yielded the diazonium chloride (II) of I which probably has the structure $\text{N}=\text{CMe}-\text{C}=\text{N}-\text{Cl}$. It is a white powder, very stable under



ordinary conditions, and heated at 100° for 6 hrs. it loses only about 20% of its coupling power. With KCN in acid soln. it loses its coupling power and yields a yellow ppt. II couples with PhNMe_2 , PhNH_2 , β - $\text{MeC}_6\text{H}_4\text{NH}_2$, α - $\text{C}_6\text{H}_4\text{NH}_2$, and β - $\text{C}_6\text{H}_4\text{NH}_2$ (orange-yellow crystals, m $195-6^\circ$) to give colored azo derivs., with β - $\text{C}_6\text{H}_4\text{OH}$ it gives a red compd. m $180-90^\circ$, with 8,3-b- and 2,6-S- $\text{C}_6\text{H}_3(\text{OH})(\text{SO}_3\text{H})_2$ it gives sol. red dyes; with 1,3,3,6-C₄H₄(OH)₂(SO₃H)₂ it gives a sol. purple dye, and with m - $\text{C}_6\text{H}_4(\text{NH}_2)_2$ it gives a red azo chrysoidine color. It also couples with AgClH_2 to give dark yellow crystals, m $84-5^\circ$, and with $\text{AcCH}_2\text{CO}_2\text{Et}$ to give yellow crystals, m $115-6^\circ$. These compds. with diketones are sol. in the usual org. solvents but not in alkalies and may possess one of the following 3 structures (Pyr = pyrazole nucleus) Pyr-N-N-CH(COR')COR, Pyr-N.NC(COR')C(OH)R, Pyr-NHN-C(COR')COR.

C. H. PERR

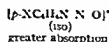
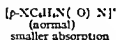
Organic catalysts. V. Investigation of the cleavage of α -keto acids. WOLFGANG LANGENBECH, RUDOLF HETSCHENREUTER AND ROBERT JETTEMANN, Ann 455, 53-61 (1931); cf. C. A. 24, 42N.—The following study was made in an effort to det. the cause of the retarding action of aldehydes upon the catalytic cleavage of α -keto acids. BrCO_2H , PhNH_2 (2 g.) and 0.1 g. BrCO_2H (I) or 5 drops BzH , heated at 250° for 5 min., give 0.05 g. of 1-phenyl-5,6-dihydro- α' , β' -naphthacarbazole (II), $\text{C}_{16}\text{H}_{15}\text{N}$, pale yellow, decomp. between $400-50^\circ$, a part of the compd. being sublimed; it crystallizes with 2 mols. H_2O , which are not lost at 150° in vacuo. It also results in 0.5 g. yield from 5 g. desyl chloride

and 2 g PhNH₂ at 140–50° for 1 hr., removing the PhNH₂, HCl and again heating with BzI at 150–60° for 1 hr. PhC(NPh)CHPhNHPh is probably an intermediate product, since with BzI at 150–60° for 1 hr., this yields 72% of II. Perchlorate of II, needles from boiling AcOH, decomps when heated; peroxide, brown leaflets, becoming colorless at 250°, losing I; perbromide, needles, HCl sol., needles, decomps above 300°, with KI this gives the difficultly sol III salt. II does not give a crystalline oxidation product with O₂ in AcOH. CrO₃-AcOH gives a chromate but has no oxidizing action. II is probably formed in a similar manner in the catalytic cleavage of I by PhNH₂; β-Aminooxindole forms a perate (III), yellow, darkens at 150°. Further expts on the action of PhNHCH₂CO₂H, BzCO₂K, p-MeOC₆H₄NHCH₂CO₂H, m-O₂NC₆H₄NHCH₂CO₂H and III on I are reported. C. J. West

Absorption spectra and constitution of azoxy derivatives and analogous compounds. L. SZCZÓ AND P. OSTINELLI. *Atti III congresso naz chim pura applicata* 1930, 335–401.—The ultra violet absorption spectra of alc solns of desoxybenzoin and the following derivs have been detd: α-p-nitro, α-p-amino, α- and β-p-hydroxy, α- and β-p-methyl and β-p-chloro (α form is C₆H₄COCH₂CH₂Ph, β form is C₆H₄CH₂COCH₂Ph). β-p-Chloro desoxybenzoin is prepd as follows: phenylacetyl chloride is condensed with chlorobenzene (I) in the presence of AlCl₃ by heating 2 hrs. on a water bath and 2 hrs. at 120°. The excess of (I) is distd off and the mass decompd with H₂O and then extd. with ether. This was distd off and the residue treated with animal charcoal in 95% alc. It was pptd with H₂O and recrystd from ether; m 102°. The oxime was prepd by heating an alc. soln of the ketone with the equiv quantity of hydroxylamine-HCl and Na₂CO₃ in H₂O. The alc. was evapd and the product, washed with H₂O dissolved in alc. and pptd with H₂O, m 117°. The α isomers and non-substituted desoxybenzoin behave alike but give a different type of absorption curve from the β-isomers. J. B. A.

Absorption spectra and the constitution of the desoxybenzoin series. L. SZCZÓ AND P. OSTINELLI. *Gazz chim ital* 60, 677–88 (1930).—In a previous paper (cf. preceding abstr.) it was shown that substituents have different effects on the absorption spectra of desoxybenzoin compds from their effect on those of azoxy compds (cf. S. C. A. 23, 4406). In the latter the influence of the substituent in the nucleus on the absorption power is less when it is nearer the azoxy group (XC₆H₄N(O)NPh) than when it is in the farther position (XC₆H₄N(O)Ph). On the contrary, absorption is greater in XC₆H₄COCH₂Ph compds than in XC₆H₄CH₂COPh compds. The behavior of the azoxy compds is explained by the fact that double bonded O in the azoxy group dampens the effect of the substituent by partial valences, resulting in a more stable condition in the substituted nucleus. This hypothesis is supported by the great similarity between azo and XC₆H₄N(O)NPh compds. The opposite effect in the desoxybenzoin series does not permit a similar hypothesis, so further expts were carried out with isomers of other desoxybenzoin (substituted in both nuclei), most of which are new. p-ClC₆H₄CH₂CO₂H cannot be prepd satisfactorily by the method of Petrenko-Kritschenko (cf. Ber. 25, 1240 (1832)), so it was prepd. by sapon of p-ClC₆H₄CH₂CN. It was then transformed into p-ClC₆H₄CH₂CO₂Cl (I). The latter condensed with C₆H₅ in the presence of AlCl₃ yields p-ClC₆H₄CH₂COPh (II). p-MeC₆H₄CH₂Cl condensed with PhCl in the presence of AlCl₃, heated several hrs. at 130°, excess PhCl distd off, the residue decompd with water, washed with HCl, then with Na₂CO₃, extd. with Et₂O, the residue dissolved in EtOH, purified with animal charcoal, pptd with water and crystd repeatedly from a mixt. of petr ether and C₆H₆, yields 4-chloro-4'-methyl-desoxybenzoin, p-MeC₆H₄CH₂COC₆H₄Cl p (III), m. 123°. Similarly, condensation of I with PhMe yields 4-methyl-4'-chloro-desoxybenzoin (IV), m 113°. MeC₆H₄CH₂CO₂H, PhOH and ZnCl₂, heated 0.5 hr. at 200°, dissolved in C₆H₆ and a little water, washed free of ZnCl₂ by HCl, made alk. with Na₂CO₃, acidified with HCl (d 1.19), the ppt purified with animal charcoal and recrystd from dil. EtOH, yields 4-hydroxy-4'-methyl-desoxybenzoin (V), m. 162°. p-O₂NC₆H₄CH₂Cl, PhMe in CS₂ and AlCl₃, heated 3 hrs. on a water bath, the CS₂ and PhMe distd., the residue decompd by water washed with HCl and Na₂CO₃, crystd from dil. EtOH, reduced with Sn and HCl in EtOH, made alk., extd. with Et₂O, the ext. evapd, the residue recrystd from dil. EtOH, diazotized in H₂SO₄ and the product recrystd from hot water, yields 4-methyl-4'-hydroxy-desoxybenzoin (VI), m 146°. 4-Hydroxy-3-methyl-desoxybenzoin (VII) was prepd by the method of Blau (*Monatsh* 26, 1149 (1905)). PhCH₂Ac was prepd by dry-distg a mixt. of Ca(OAc)₂ and (PhCH₂CO)₂Ca (2 parts). The measurements of the absorption consists in the ultra violet range were carried out by the method already described (*loc cit*). The ketones were examd in hexane (optically pure) (cf. Wolf and Herold, C. A. 23, 5416) and the desoxybenzoin in EtOH (the absorption curves with EtOH and with hexane are identical). The results are tabulated

in complete detail. In all cases the phenomena were similar to those previously described, viz., the max absorptions of the *d* isomers of *p*-chloro-, *p*-methyl- and *p*-hydroxydesoxybenzoin are displaced toward lower frequencies with respect to their *a* isomers. The displacement between the 2 OH isomers is greater than between the 2 Cl derivs. which in turn show greater differences in the isomers with a nuclear Me group. The weak absorption band of the *d* isomers of hydroxy-, methyl- and chlorodesoxybenzoin, and to which has been attributed the color, disappears after complete purification. Thus, *d*-hydroxydesoxybenzoin, described by Weisl (*Monatsh.* 26, 19-1 (1905)) as orange red, is colorless when highly purified. In the compds with substituents in each nucleus, as in the methylhydroxy derivs., V has a great max. displacement compared with that of VI and the displacement increases still more when both substituents are in the nucleus on the CO group, viz., in VII. Similar relations are found in the Me Cl derivs., where the effect of the Cl prevails over that of the Me group, as with the compds contg. a single substituent. Therefore the increase in light absorption resulting from proximity to the CO group is not unusual but is found in all compds. The CO group has an effect wholly apart from a dampening of the substituent and its nucleus. The proximity of CO to unsatd. bonds greatly increases the absorption power (cf. Bielecki and Henni, *C. A.* 8, 691), and even the C_{6H}_5 nucleus with its conjugated double bonds has a similar influence (cf. v. Auwers, *C. I.* 9, 14)¹⁰ 13, 432-3). Measurements of ketones of the $PhCOEt$ and $PhCH_2COMe$ type show how profound is the influence of the position of CO with respect to the C_{6H}_5 nucleus. In every case there is an increase in the power of extinction, beginning with compds with the CO directly on the C_{6H}_5 nucleus. This effect is much more intense in the ketones than in the desoxybenzoins, since in the ketones the distance between 2 centers of oscillation has a greater influence than in the desoxybenzoins, in which the displacement of frequency indicates only the variations in the state of satn. of 1 C_{6H}_5 nucleus with respect to the other. The joint effect of the 2 centers of oscillation with unsatd. character leads to greater optical exaltation. In desoxybenzoins the CO group has a lower degree of satn. than the neighboring CH_2 group, while the azoxy $-N(O)=$ group corresponds to nitric N and involves a greater satn. of the same N atom. Therefore, of the 2 N atoms in the azoxy compds., the quinquivalent N is in the higher state of satn. The proximity of an O or quinquivalent N atom should not have a depressing effect on the substituted nucleus, while the trivalent N has an exaltant effect like the CO group, which induces an increase in the absorption power by the reciprocal influence of the 2 unsatd. groups. A comparison of the absorption spectra of the simplest aromatic derivs. having their N in the form of NO_2 , NO or NH_2 groups confirms these deductions. The same reasoning may be extended to diazo compds. According to Angeli (*C. A.* 22, 2991) these are not simply stereoisomers. Cambi (*C. A.* 21, 3895) on the contrary gives the structures $(RN(O):N)^+K^+$ and $(RN:N:O)^+K^+$ to the normal and isodiazotates, resp. Cambi and Szegő (*C. A.* 23, 337) have shown that normal diazotates have continuous absorption spectra with low coeffs. of extinction, whereas the iso derivs. have a max. absorption in the ultra-violet. Based on the formula above of normal diazotates, the quinquivalent N carrying the O is next to the nucleus, while in iso compds. the O is on the terminal N. The optical behavior agrees perfectly with the hypothesis



Both in the aromatic ketones with the CH_2CO chain and in azoxy derivs. and diazotates, the variation in absorption power for ultra violet light depends upon the distance of the 2 centers of oscillation with the lower degree of satn., and it increases when the 2 centers are closer together.

C. C. DAVIS

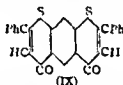
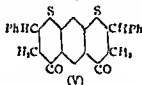
Spectrography of the flavone series. III. The constitution of wogonin. SHIZUO HATTORI. *Acta Phytochim.* 5, 99-116(1930), *cf. C. A.* 22, 3661—Wogonin (I) is present in small amounts in the roots of *Scutellaria baicalensis* Georgi (*Labiatae*) while baicalin forms the much larger part. I is extd. with benzene and recrystd. from 80% alc., forming yellow needles, m. 201°. The analyses correspond with the formula $C_{15}H_{10}O_5$; it contains two HO and one MeO group. One of the HO groups is methylated by CH_3N_3 , forming a Me ether, m. 183°, which gives a brown color with $FeCl_3$. This behavior indicates a free HO group in the *o*-position to the C:O group. Methylation with Me_2SO , and alkali yields a di-Me ether, m. 167-8° (II), which does not give a color with $FeCl_3$. Upon hydrolysis with 20% NaOH, $PhCOMe$ and $BzOH$ are obtained. Upon boiling with HI (d. 1.7), I is demethylated with the formation of a deep yellow

product, m 225-7°, which has three HO groups, probably in the 5,7,8-positions, and the formula $C_{15}H_{12}O_6$. Comparison of the absorption spectra of acetilwogonin with 8-hydroxy- and 4-methoxyflavone showed an analogy from which it is concluded that I has the MeO group in position 8. The structure was finally detd. by synthesizing 5,7,8-trimethoxyflavone and comparing it directly with II. The reaction of $AlCl_3$ with 2,5,4'- $(MeO)_3C_{15}H_{10}O_6$ yielded a substance, m 112-3° (III), identical with I of the 4 substances obtained by Bazzelloni and later by Chapman, Perkin and Robinson (C A 22, 341) contg. one HO and three MeO groups. The HO group gave a FeCl₃ reaction indicative of being in the 8-position to the C O group (2 or 6). Fusion of III with Br₂O and H_2O yielded a flavone, m 167-8°, identical with II and not with the known 5,6,7-trimethoxyflavone I, therefore, is 5,7-dihydroxy 8-methoxyflavone. Absorption spectra are given of 8-methoxyflavone, acetil-, mono- and dimethylwogonin and wogonin and 5,7,8-trihydroxyflavone. I. GERHART FRIEDTGG

Flavone group. V. m-Methoxybenzaldehyde and diacetoresorcinol. ILIA BETH M. KLAN AND HUGH REYN. *Proc Roy Irish Acad* 30B, 425-33 (1930), cf C A 8, 1107, 9, 2000 10, 1700, 1949-9, 2030, 2037. —The diacetate of tetrabromodibenzylidenediacetoresorcinol (I) yielded diflavone upon treatment with alc. KOH, but the corresponding anisylidene and vetatylidene compds. yielded dicoumaranones. The type of ring formed must depend upon the relative labilty of the α - and β -Br atoms and it appeared that a β -MeO group increased the labilty of the β -Br atoms. According to the theory of induced alternate polarities, a m-MeO group should have a less marked influence so the m-MeO homolog (II) of I was prepd. by condensing diacetoresorcinol with m $MeOC_6H_4CHO$, followed by acetylation and bromination. The principal product was the corresponding 3',3'-dimethoxydicoumaranone (III) but the absence of a diflavone was not proved. Di-m-methoxybenzylidenediacetoresorcinol di Me ether slightly sol in EtOH Et₂O, sol in PhH, Me₂CO $CHCl_3$, m 157-8°, gave a dibromide m 213° (decompr.), a tetrabromide (purity uncertain), m 176° (decompr.), and a pentabromide (purity uncertain), m 18° (decompr.). Di-m-methoxybenzylidenediacetoresorcinol (IV) was mol in H₂O and petroleum ether, slightly sol in EtOH and Et₂O < in dil KOH, $CHCl_3$, Me₂CO and PhH and m 157°. IV prepd. in EtOH was orange-yellow, better yields were obtained in MeOH but the color was light yellow. The resp. colors of these chromonomers remained unchanged after repeated recrystns. their m ps and mixed m ps. were the same and they formed the same diacetate m 122°. II could not be crystd. III, crystd. from C_6H_6N , m 277-8° (decompr.). Dipiperonylidene-m-methoxydiflavone prepd. by treating a mixt. of 2 g of IV and 5 g of piperonal in 60 cc. of $CHCl_3$ with 4 cc. alc. HCl m 245°. Similarly BzH condensed with IV to give dibenzylidenedi-m-methoxyflavone, m 249-50°. C H PIER

Thiophenols. Dithioflavone and linear dithioflavonone. C FRYL G VENTURENT AND L. SARTINI. *Gazz chim ital* 63, 79-81 (1933). —Previous expts. which dealt with compds. contg. 2 heterocyclic sulfurated nuclei arranged symmetrically in the C_6H_4 nucleus and in which dithiothiazine, dithiocoumarone and dithioxanthone were prepd. (cf C 4 20, 1992 21, 730 22 41) were continued in the present work by the prepn. of dithio derivs. It was necessary in this connection to study the behavior of polythiophenols with unsatd. compds. of the C_6H_4 and C_6H_2 series. Dithioresorcinol (I) was chosen because of the ease of prepn. of its isomers and because from it could be obtained derivs. contg. 2 sym. heterocyclic sulfurated groups on the central C_6H_4 nucleus. It was found that I forms addn. products with C_6H_4 and C_6H_2 derivs., each group adding to a double or to a triple bond. The addn. of PhCH $CHCO_2H$ (II) to I can be accomplished by heating I, II (2 mols.) and glacial AcOH satd. with HBr at 100° in a sealed tube for several hrs., steam-distg. (to remove excess I) and washing the residue with boiling water, but the yield is very small. A much higher yield is obtained by heating I, PhCH $CHCO_2H$ (cf Ber 11, 1231) (3.5 parts) and AcOH satd. with HBr (1 part) at 100° in a sealed tube for 5-6 hrs., pouring into water, steam-distg. and washing the residue with boiling water. The product is m-phenylenedithio- β -acryl- α -cyanamic acid m $C_6H_4[SCH_2CH(CO_2H)Ph]_2$ (III) (pale yellow) which cannot be crystd., sol in aq. alk. hydroxides and carbonates. III (5 g.) in aq. Na_2CO_3 oxidized with aq. $KMnO_4$, 2 g. the reaction mixt. being partially neutralized with dil H_2SO_4 , during the oxidation) filtered, washed with boiling water, the filtrate concd., acidified with H_2SO_4 , and the ppt. recrystd. from EtOH, yields m-phenylenedisulfonyl- β -hydroxycyanamic acid, m- $C_6H_4SO_2CH_2CH(CO_2H)Ph_2$ (IV), m 192°. III and $POCl_3$ (15 parts), heated on a water bath for a long time (with vigorous agitation), poured into ice-water (before the reaction product has become too thick), the sep'd. green-brown oil allowed to solidify, digested with warm dil. Na_2CO_3 extd. with boiling EtOH, the ext. concd., and the ppt.

purified by repeated crystals from 1 tOH, yield *dithioflavanone* (V), greenish yellow, softens around 65–70°, semifuses at 85–90°, completely fused at 129.5°, gives intense carmine red solns in concd H_2SO_4 (test of thiocinnamyl group). IV is also formed from III and PCl_5 with anhyd AlCl_3 under the same conditions. I (5 g) in anhyd PhMe (40 cc) and Na (2 g), boiled several hrs., with frequent agitation, the unaltered Na carefully removed, PhC CCO_2Et (12.5 g) in PhMe added, with vigorous agitation, let stand several hrs., water added, acidified with dil H_2SO_4 , extd with Et_2O , the PhMe Et_2O extd with aq KOH, the PhMe Et_2O residual soln dried with CaCl_2 , and the solvents distilled *in vacuo*, leaves a residue of straw yellow oil which is a mixt of *Li m*-phenylene α -dithiocinnamate and *Li m*-phenylene β -dithiocinnamate which could not be crystal. This mixt., saponified by refluxing with alc. KOH (4 g) for 2 hrs., dild with water (equal vol), the 1 tOH distd, the aq residue extd with Et_2O (to remove traces of styrene formed by decompn of the acid), the aq residue acidified with dil H_2SO_4 , the ppt (mixt of α and β acids) extd with warm 1 tOH, and the ext. evapd., leaves *m*-phenylene α -dithiocinnamic acid, $m\text{-C}_6\text{H}_4[\text{SC}(\text{CH}_2\text{Ph})\text{CO}_2\text{H}]_2$ (VI), oil. The residue from the 1 tOH extn., recrystd from boiling AcOH, yields *m*-phenylene β -dithiocinnamic acid, $m\text{-C}_6\text{H}_4[\text{SC}(\text{CH}(\text{CO}_2\text{H})\text{Ph})_2]$ (VII), pale yellowish, m 221°. VII heated above its m p. evolves CO_2 and leaves *m*-phenylenedithiostyrene, $m\text{-C}_6\text{H}_4[\text{SC}(\text{CH}_2)\text{Ph}]_2$ (VIII), oil, cannot be crystal. Also I (4 g) in anhyd PhMe and Na (1 g), refluxed for a long time, cooled, the excess Na removed, agitated with HC CPh (4.5 g) in PhMe, let stand, a little water added, acidified with dil H_2SO_4 , extd with Et_2O , the PhMe Et_2O extd with aq NaOH, the residue washed with water, dried with CaCl_2 , and the Et_2O and PhMe distd., leaves VIII. Na Hg (90 g of 3%) added slowly to VII (1.5 g) suspended in water, acidified with dil H_2SO_4 , forms an oil which cannot be crystal. This oil (1 g) in aq Na_2CO_3 oxidized with concd KMnO_4 (calcd quantity), filtered, washed with boiling water, the filtrate concd., acidified with dil H_2SO_4 , and the ppt recrystd from boiling 1 tOH, yields IV. PCl_5 (10.5 g) and then anhyd AlCl_3 (4.5 g) added with continuous agitation to VII (2 g) suspended in anhyd CaH_2 (20 cc), warmed gently to complete the reaction, let stand with frequent agitation, cooled, ice added, filtered, the residue washed with water, digested in dil aq Na_2CO_3 , and the residue recrystd twice from boiling CaH_2 , yields *dithioflavanone* (IX), pale lemon-yellow, m 255.6°, gives a yellow soln in concd H_2SO_4 . VII (1 g), agitated with concd



H_2SO_4 (5 cc) (cooling), heated at 50–60° for some time, poured into water, filtered, the residue washed with water, digested in aq Na_2CO_3 and recrystd from CaH_2 , also yields IX. VI in oil form (*loc cit*) could not be made to crystal from any solvent, but by acidification of its soln in aq Na_2CO_3 it seps in the solid state, though not crystal. The only means of purification is to dissolve it in dil Na_2CO_3 , filter, reprecipitate with dil H_2SO_4 , and repeat this operation. Thus prepd., VI is a yellowish amorphous solid, which in time becomes compact and friable. It has no sharp m p., but softens around 70° and at 73° it becomes a frothy liquid. Hydrogenated with Na-Hg and then oxidized with KMnO_4 , VI partially decomps., but does not yield crystallizable products. It is decomposed by concd H_2SO_4 , with evolution of SO_2 and H_2S , in this case the H_2SO_4 not having the dehydrating action which occurs in the formation of the thiocinnamyl and thiocinnamyl nuclei, but rather an oxidizing action which results in cleavage of the double bond. VI, heated at 180–90° until CO_2 is no longer evolved, cooled, washed with aq Na_2CO_3 (to remove unaltered VI), suspended in 1 tOH, reduced with a great excess of Na-Hg, the Hg sepd., filtered, the filtrate partially evapd., dild with a large vol of water, let stand, yields a light yellow oil which cannot be crystal and which is *m*-phenylenedithioethylbenzene, $m\text{-C}_6\text{H}_4[\text{SC}(\text{CH}_2\text{CH}_2\text{Ph})_2]$ (X), does not distil up to 300° *in vacuo*. Aq NaOH (12 g) and then $\text{PhCH}_2\text{CH}_2\text{Cl}$ (4 g) added to alc I (2 g), refluxed 0.5 hr., the 1 tOH distd., the residue extd with Et_2O , the ext. dried with CaCl_2 , and the Et_2O evapd., yields X. This confirms the constitution of VI and VII, for in the reaction between I and PhC CCO_2Et , the addn. of the thiophenol groups on the C_6H_4 bond takes place in 2 ways: union of the S to the α -C and to the β -C atom of the PhC CCO_2H , giving rise to 2 groupings PhC \cdot $\text{CH}(\text{CO}_2\text{H})$ S— and PhCH \cdot C \cdot CO_2H S—.

C. C. DAVIS

The position of the glucuronic acid linkage in baicalin. KEIYA SHIBATA AND SHIZUO HATTORI *Acta Phytochem* 5, 117-8 (1930).—In baicalin (I), a glucuronic acid deriv. of baicalin (5,6,7 trihydroxyflavone), the glucuronic acid radical (II) can be linked only in positions 6 or 7 since on methylation only 1 mono-Me ether can be obtained which gives a violet brown color reaction with FeCl_3 , showing that the HO group in position 5 and ortho to the C O group is still free. Position 7 for the linkage with II is likely on account of the green color reaction of I with FeCl_3 in analogy with the 5,6-dihydroxyflavone. Furthermore cobalt pentammine salts, which are known to oxidize vicinal phenolic derivs., attack alc. solns of I easily, forming ppts. after a color change. E. GEBAUER FUELNEGG

Quinoline compounds containing arsenic. II. Synthesis of 6-methoxyquinoline derivatives of aminophenylarsonic acids by the use of 4-bromo-6-methoxy-2-methylquinoline. ROBERT H. SLATER. *J. Chem. Soc.* 1931, 107-18, cf. *C. A.* 24, 4039.—An improved method is given for the prepn of 4-hydroxy-6-methoxy-2-methylquinoline (0.1% yield), POBr_3 gives the 4-Br deriv. (I), m. 117-8°, the AcOH and dil. HNO_3 solns. exhibit a blue fluorescence and a dil. H_2SO_4 soln. a greenish blue fluorescence, changing to violet on diln. Heating a mixt. of 8.5 g. I, 5.4 g. o- $\text{H}_2\text{N.C}_6\text{H}_4\text{AsO}_2\text{H}$, 70 cc. AmOH, 16 g. K_2CO_3 and a trace of Cu bronze and I at 130-40° for 22 hrs. gives 5.9 g. o-6'-methoxy-2'-methyl-4'-quinolylaminophenylarsonic acid (II), m. 302-3° (decompn.), isoelectric pt. pH 7.3, crystals from dil. AcOH. A colloidal soln. does not give a color with 0.001 N I, addn. of a few drops of concd. HCl gives a brilliant violet color, which disappears on addn. of more HCl , the warm concd. H_2SO_4 soln. exhibits a blue fluorescence. Various salts are qual. described. Reduction of II in EtOH HCl contg. a trace of I with SO_2 gives 12-chloro-7-methoxy-11-methyl-5,12-dihydroquinobenzarsazine (III), pale yellow, m. 245-7° (decompn.). Oxidation of III with H_2O_2 in AcOH gives 7-methoxy-11-methylquinobenzarsazinic acid, pale yellow, does not m. 310°, various solns. exhibit a blue fluorescence; salts are qualitatively described. PCl_5 below 80° gives the chloride, pale yellow, m. 165-7°, the dil. aq. NaOH soln. exhibits a bright blue fluorescence, 1 (5 g.) and o-tolidine (10 g.), heated 9 hrs. at 140-50°, give 7.2 g. of 4-o-tolidino-6-methoxy-2-methoxyquinoline (IV), pale yellow, crystg. with EtOH, m. 182°, heating 5 hrs. at 100° gives the anhyd. compd., m. 199-200°, the diazo chloride gives a blood red azo dye with alk. β - $\text{C}_{10}\text{H}_7\text{OH}$, a dil. soln. in EtOH or AcOH gives a reddish brown color with 0.001 N I. Ac deriv., m. 182-3°. The diazo compd. from IV, pale orange, with Na_2AsO_3 gives 4',6'-methoxy-2'-methyl-4'-quinolylamino-3,3'-dimethyldiphenylarsonic acid, m. 304-5°; the warm concd. H_2SO_4 soln. shows a light blue fluorescence, which disappears on diln. 4-o-Dianilino-6-methoxy-2-methylquinoline (V), pale yellow, m. 195-6°, was best prepd. by heating the components at 140° and 10-5 mm., a dil. EtOH soln. gives a red color with 0.001 N I. Ac deriv., crystals with H_2O , m. 140° and then 200-1°, the C_6H_5 soln. exhibits a greenish blue fluorescence, the EtOH soln. gives a brilliant blue color with 0.001 N I, while the AcOH soln. gives an intense red color. The Bart reaction with V gives 4',6'-methoxy-2'-methyl-4'-quinolylamino-3,3'-dimethoxydiphenylarsonic acid, pale grey, m. 243-5° (decompn.), the dil. AcOH soln. gives a dark brown ppt. with 0.001 N I. I and (p- $\text{H}_2\text{N.C}_6\text{H}_4$) CH_3 , heated at 135-40° and 10-5 mm., give p-6-methoxy-2-methyl-4-quinolylamino-p'-aminodiphenylmethane, pale yellow, crystals with 1 mol. H_2O , m. 135-45°, the Bart reaction gives p-6-methoxy-2-methyl-4-quinolylaminodiphenylmethane p'-arsonic acid, pale yellow, chars at 300°, a soln. in concd. H_2SO_4 gives a brilliant purplish blue fluorescence which is destroyed on diln. with H_2O . 4-Anilino-6-methoxy-2-methylquinoline, m. 208-9°, HCl salt does not m. 310°. The chemotherapeutic tests will be published elsewhere. C. J. WEST

β -Diketones in ring formation. II. UMAPRASANN BASU. *J. Indian Chem. Soc.* 7, 810-24 (1930) cf. *C. A.* 24, 5752.— Bz_2CH_2 (I), supposedly 96% enolic, gives 45% of 3-cyano-4,6-diphenyl-2-pyridone, when condensed with $\text{CNCH}_2\text{CONH}_2$ (IV). To exam. the influence of the adjacent substituent in the enolic modification in the course of such condensations, IV has been condensed with 1, p-MeC $_6\text{H}_4\text{COCH}_2\text{Bz}$ (II) and $\text{CH}_2(\text{COEt})_2$ (III). I condensed with IV in the presence of NaOEt , gave 5-20% of

$\text{NH.CO.C(CN).CPh.CH.CPh}$, m. 320°. By using Et_3NH as a condensing agent, 55-70% was obtained, a 50% excess of IV improving the yield. The condensation

of IV with II by the aid of Et_3NH gave 34% of $\text{NH.CO.C(CN).CPh.CH.CC}_6\text{H}_4\text{Me}$, m. 267° (which on heating with H_2SO_4 evolved CO_2 and formed

$\text{NH.CO.CH.CPh.CH.CC}_6\text{H}_4\text{Me}$, m. 226-8°), and 17% of NH.CO.C(CN).C-

$(C_6H_5)_2CH\ CPh$, m 311-2°, hydrolyzed to $NH\ CO\ CH\ C(C_6H_5)_2CH\ CPh$, m 237-9°. According to B the main crop is derived from the $PhC(OH)CHCO_2C_6H_5$ modification owing to the greater inhibiting effect of the β tolyl group in the $MeC_6H_4C(OH)CHCO_2Ph$ form. That the reaction had taken place only through

the enolic phase was evident by the isolation of only $NH\ CO\ C(CN)CPh\ CH\ CC_6H_5$ from the condensation of IV with $PhC(OMe)CHCO_2C_6H_5$.

$NH\ CO\ C(CN)C(C_6H_5)_2CH\ CCl_2$, m 184°, the condensation product of III and IV, was heated with concd HCl for 3 hrs at 150° and hydrolyzed to $NH\ CO\ CH\ CH_2CH\ CH_2$, m 61.2°. The influence of substituents on the reactivity of the CH_2 group of IV was studied by condensations involving the use of $CH_2(CONH_2)_2$ (VI) and $CH_2(CN)_2$ (VII).

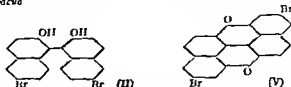
Ac_2CH_2 (VIII) and VI formed $NH\ CO\ C(CONH_2)_2CMeCH\ CMe$, m 224.5° (which was hydrolyzed with concd HCl to β -lutidostyrol, m 179.81°), whereas $CICH_2CONH_2$ and $PhCH_2CONH_2$ could not be condensed with VIII. A mixt of VII and VIII with the addn of a little Li_2NH gave 77% of 3 cyano- β -lutidostyrol,

$NH\ CO\ C(CN)CMeCH\ CMe$, m 280°. With $CNCH_2CO_2Et$ and VII, resp. $BrCH_2I_2$ gave $NH\ CO\ C(CO_2Et)CMeCH\ CPh$, m 216-7°, and $NH\ CO\ C(CN)CMeCH\ CPh$, m 310°.

C. R. ADDONALL

Derivatives of dinaphthylene dioxide. A CORBELLINI AND PASTRINI *Gazz. chim. ital.* 60, 843-61 (1930). —The expts continue the previous ones (cf. C and Vietti, C. A., 23, 3458). Concd β -dinaphthol (I) in $AcOH$ and Br (calcd quantity) in $AcOH$, heated on a water bath, cooled, the ppt. removed, the mother liquor concd until further crystn is complete, and the combined products recrystd from C_6H_6 or $CHCl_3$, yield almost 100% of dibromo- β -dinaphthol (II), m 202-3° (which is far from the 155° obtained by Fosse (cf. *Bull. soc. chim.* [3], 21, 657), when prepd in the same way). Methylation of alk. II by Me_2SO_4 with crystn from $PhCl$ yields the di-Me deriv., $C_{10}H_6O_2Br_2$, m 239.5° (not cor.). To det. the position of the 2 Br atoms the oxidation products of II were studied. With a compd of this complexity, the products vary with the nature and quantity of the oxidizing agent. II (15 g) in $AcOH$, oxidized with CrO_3 (15 g) on a water bath, when the reaction is completed heated to boiling, concd, dil'd with water, acidified with H_2SO_4 , the ppt. dissolved in dil. Na_2CO_3 , oxidized with $KMnO_4$ (at first cold, then on a water bath), the excess $KMnO_4$ reduced with $EtOH$, filtered, the filtrate concd. on a water bath, acidified with HCl , extd. with Li_2O , the ext. evapd almost to dryness, $CHCl_3$ added, the ppt. sublimed *in vacuo* around 150° and the sublimate recrystd. from abs. Et_2O , yields 4 bromophthalic acid (III), m 105-7° (cf. *Ber.* 20, 1017 (1887), Stephens, C. A. 16, 233). Oxidation in the same way with $KMnO_4$ of the product obtained by extg with Et_2O the dil'd $AcOH$ soln., or that obtained by sepn. on long standing, yields more III. Alk. II (5 g), oxidized with dil. aq. $KMnO_4$ (835 g) first cold and then on a water bath, the excess $KMnO_4$ reduced with $EtOH$, filtered hot, the residue washed thoroughly with boiling water (to ext. the slightly sol. K salt), the combined filtrates acidified with HCl and the ppt. recrystd. from $PhNO_2$, yields 2-[2-hydroxy-6-bromo-1-naphthoyl]-5-bromobenzoic acid (IV), pale brown-yellow, m 346° (not cor.). The Na salt of IV (0.2 g) and fused KOH (1 g), kept at a min. temp. until the product is completely sol. in water, dissolved in water and aq. HCl added, ppt. an unidentified compd., does not m. up to 300°, contains no Br , sol. in aq. Na_2CO_3 . This reaction with KOH was carried out with the purpose of decomp. IV into III and bromo-2-naphthol, a decompn. which takes place easily with 2-[2-hydroxynaphthoyl]benzoic acid (cf. Walder, *Ber.* 16, 299 (1883)), but in the present case KOH probably merely replaced 2 Br atoms by OH groups, without rupture of the mol. Boiling aq. 6-bromo-2-naphthol (m 128.5°, cf. Franzen and Stauble, C. A. 16, 1945) (0.25 g in 150 cc.), oxidized with boiling aq. $FeCl_3$ (0.3 g), boiled for some time, filtered, the residue washed with dil. HCl and recrystd. from C_6H_6 , yields 2,2'-dihydroxy 6,6'-dibromo-1,1'-dinaphthyl, m 203°, mixed with II it melts at the same temp., establishing the constitution of II. II cyclized by CuO according to the method already used for I (cf. C. and Vietti, *loc. cit.*), the product extd. with $PhCl$, the ext. concd. and the ppt. sepd. by fractional crystn. from $PhNO_2$, yields 2 products: (1) a less sol. compd., $C_{10}H_6O_2Br_2$ (V), m. 361-2° (not cor.), yellow, corre-

sponds in compn to a di Br deriv of dinaphthylene dioxide, sublimes unaltered *in vacuo*, and (2) a more sol unidentified compd $C_{20}H_{10}O_2Br_2$, red, m 421-2°, sublimes unaltered *in vacuo*



C. C. DAVIS

Quinoline compounds. II. Some derivatives of 4-phenyl-2-methylquinoline. UPENDRA N BRAHMACHARY AND TARAPADA BHATTACHARYYA, *J Indian Chem Soc* 7, 831-4 (1930) cf C A 24, 5875.—The enhanced toxicity of 4 phenyl 2-methylquinoline toward protozoa compared to that of quinoline induced B and B to synthesize 6-nitro-4-phenyl-2-methylquinoline (I) and 8-nitro-4-phenyl 2-methylquinoline (II). A mixt. of 50 cc. AcPh and 30 cc. (AcH)₂, cooled in ice and sat'd with HCl was allowed to stand for 48 hrs. After resatin with HCl, and standing for 24 hrs the mixt. was added to 58 g. of $p\text{-NO}_2\text{C}_6\text{H}_4\text{NH}_2$ and 90 cc. conc'd HCl and the whole was heated for 5 to 6 hrs on a water bath. After diln and filtration, the cooled filtrate was freed from $p\text{-NO}_2\text{C}_6\text{H}_4\text{NH}_2$ by neutralization with NaOH and filtration. The filtrate was supersat'd with NaOH with pptn. of I, m 141°. I was reduced by SnCl₂ and HCl to 6-amino-4-phenyl 2-methylquinoline, m 188°, and condensed with $p\text{-Me}_2\text{NC}_6\text{H}_4\text{CHO}$ to 6-nitro-4-phenyl-2- p -dimethylaminostyrylquinoline, m. 64°. II, similarly prepd from $p\text{-NO}_2\text{C}_6\text{H}_4\text{NH}_2$, m. 94°, was reduced to 8-amino-4-phenyl 2-methylquinoline and condensed with $p\text{-Me}_2\text{NC}_6\text{H}_4\text{CHO}$ to give 8-nitro-4-phenyl 2- p -dimethylaminostyrylquinoline, m 129°.

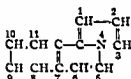
C. R. ADDONALL

Quinoline derivatives. VII. HANVS JOHN *Ber* 63B, 2637-61 (1930), cf C A 21, 914, 25, 953.—The need for prep'g large amounts of quinine acid (6-methoxyquinoline-4-carboxylic acid) (I) from quinine or conquinine (II) by oxidation with CrO₃ led to a study of the conditions for obtaining the max yield in the min. time. The expts., carried out under identical conditions as regards quality of the starting materials and reagents and the method of procedure, led to the following results. II by the Skraup method gave 42% I. Addn. of Fe (as FeSO₄) raised the yield to 53%. Co, Ni and Vd were without effect. Mn, as MnO₂, MnSO₄ or KMnO₄, at a concn. of 0.1 atom metal per l., gave 75% of very pure I; to obtain this yield the wt. of CrO₃ used must be at least twice that of the II. The optimum H₂SO₄ concn is 8%. Replacement of the Mn by Mg, Al, Cu, Ag, Ce and Pb did not raise the yield above 48%, and Hg lowered it to 28%. Cl, As and Kieselguhr also did not prove suitable. The procedure for oxidizing 100 g. II in the presence of Mn is given in detail.

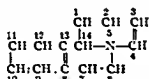
C. A. R.

1- α -Halogenoalkylquinolines and their derivatives. REGINALD CHILD AND FRANK L. PYMAN, *J Chem Soc* 1931, 36-49 cf C A 24, 119.—This work was undertaken to prep substances having antiseptical properties and followed clues indicated by the structure of emetine. The following acyl phenylethylamides were prepd as intermediates from the acid chloride and the amine in Et₂O. Chloroaceto- β -phenylethylamide (I) m. 67° (all m. p. cor.) m MeO deriv (II) m 56-7° β -picronyl deriv (III), m. 72°, β -veratryl deriv (IV) m 96° bromoaceto- β -veratrylethylamide (V), m 115°, cyanoaceto deriv (VI) m 115° but after solidifying m 127-8°, hippuro- β -phenylethylamide (VII) m 161° β -veratryl deriv (VIII), crystals with 1 mol H₂O, m 55-95° β -chloropropiono β -veratrylethylamide (IX) m 102-3°, γ -tombutoxy deriv (X) m 65° chloroaceto deriv (XI) m 60-2°, Br deriv (XII), m. 70-2° IV (30 g.), 30 cc POCl₃ and 40 cc PhMe heated at 120-30° for 1 hr give 93.8% of 1-chloro-methyl-6,7-dimethoxy 7,4-dihydroisoquinoline HCl (XIII), canary yellow, darkens 210°, effervesces 217° picrate yellow m 195° (decompn.) XIII and KCN in H₂O give 1-chloromethyl 1-cyano-6,7-dimethoxytetrahydroisoquinoline (XIV), m 125° (decompn.), if the moist gummy ppt. of XIV is transferred to KCN in H₂O and EtOH and boiled for 0.5 hr there results 95% of the 1-cyanomethyl deriv (XV), corresponding to XIII, m. 173°, the HCl salt pale yellow, m. 205-6° and readily dissolved by H₂O, picrate, orange yellow m. 225° (decompn.) XV also results in 40% yield from VI and POCl₃ in PhMe. Reduction of XV with Na and EtOH gives 40% of 1- β -aminomethyl-6,7-dimethoxytetrahydroisoquinoline, whose picrate, yellow, crystg with 1 mol EtOH, m. 205°, the di HCl salt, cream colored, m. 276-7° (decompn.) V and P₂O₅, followed by picric acid, give 70% of 1- β -bromomethyl-6,7-dimethoxy-3,4-dihydroisoquinoline

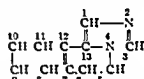
picrate, orange-yellow, m. 190–1° (decompn). V and POCl₃ give 80% of XIII. II and POCl₃, followed by picric acid, give 1-chloromethyl-6-m methoxy-3,4-dihydroisoquinoline picrate, yellow, m 169–70° (uncor), the HCl salt, m about 170°, appears to contain Me₂CO of crystn. I from III there was obtained 1-chloromethyl-6,7-methylene-dioxy-3,4-dihydroisoquinoline picrate, orange red, m 179–80° (decompn), the HCl salt decomps 204°, it also crystallizes with Me₂CO, m 150–70°. IX and POCl₃ give a mixt of 2 picrates, m 235° and 145–85° (both uncor), the 1st contained no Cl, the 2nd only a trace. Boiling X and POCl₃ in PhMe for 30 min gives a mixt of 1-γ-chloro-propyl-6,7-dimethoxy-3,4-dihydroisoquinoline picrate, which m 163–4° (not pure), and 4,13-dehydro-9,10-dimethoxy-1,2,3,5,6,13 hexahydrobenzopyrrocolinium picrate, yellow, m 201–2°, which is explosive, the chloride (XVI) crystallizes with 2H₂O, m 120–2°, or, anhyd, m 204–5°, the dry salt is hygroscopic. Reduction of XVI with Sn and HCl in EtOH gives 80–55% of 9,10-dimethoxy-1,2,3,5,6,13 hexahydrobenzopyrrocoline, m 88–9°, HBr salt, m 145°, picrate, yellow, m 187° (decompn). XI or XII with POCl₃ give over 90% 1-4-chlorobutyl-6,7-dimethoxy-3,4-dihydroisoquinoline picrate, yellow, m 156–7°, HCl salt, m 172–3° (decompn), the H₂O soln has a marked blue fluorescence. With NH₄OH the HCl salt gives 90% of 5,14-dehydro-10,11-dimethoxy-1,2,3,4,6,7 hexahydrobenzopyrrocolinium picrate, golden spangles, m 185–6°, the chloride crystallizes with 1 mol HCl and 2 mols H₂O m 197–8°, iodide, yellow, m 210–2°. Reduction gives 10,11-dimethoxy-1,2,3,4,6,7 hexahydrobenzopyrrocoline, m 59–60°, b₃ 225°, HCl salt, m 235–7°, picrate, bright yellow, m 172–4°. MeI gives a mixt. of β methiodide, m 214–5°, and the more sol α methiodide, m 228°, heating the latter at 250° gives the β salt. VIII and POCl₃ give 90% of 9,10-dimethoxy-3-phenyl-5,6-dihydrobenzoglaxalcoline (XVII), m 187°, HCl salt, m 246–7° (decompn), sol. in 25–30 parts cold H₂O, HBr salt, m 203°, methiodide, pale cream color, m 235°, picrate, yellow, m 220–7°, NO₂ deriv, orange-yellow, m 202°. Heating XVII with concd HCl 5 hrs at 100° gives the 9,10-di H₂O deriv, whose HCl salt crystallizes with 3H₂O, m. 293°, FeCl₃ gives a deep green color. XVII prevented the growth of *Entamoeba histolytica* in cultures at a diln of 1 in 25,000, whereas the control substance, emetine, was effective in a diln of 1 to 500,000. All other compds. were far less effective. The salts had no antimalarial activity and were also found to be devoid of trypanocidal activity. The nomenclature used above is based on the following types:



Benzopyrrocoline



Benzopyrrocoline



Benzoglaxalcoline

C. J. WEST

The γ-triazines. Syntheses of arylaminothioltriazines. A OSTROGOVICH AND V. GALEA. *Atti accad Lincei* 11, 1109–16(1930)—Using the method of prepn described in a previous paper (C. A. 25, 705), i. e., condensing cyanoguanidine with the arylthiol acid, 5 new arylaminothioltriazines have been prepd: phenyl (I), PhC₆H₄N C(NH₂)·N·C(SH)N, m. 281–2° (decompn); o-tolyl (II), m. 243–4° (de-

compn); the m- (III) and p-derivs. (IV), m. 272–3° and 279–80°, resp; p-anisyl (V), m. 282–3° (decompn.). The metallic salts of some of the above, as well as some esters, were prepd. I picrate, pptd from the EtOH soln by adding picric acid, m. 193–4° (decompn). I also gives insol. salts with many heavy metals as Ag, Cu, Hg, Pb, Fe, etc. The Ag salt shows very slight discoloration in sunlight. I is also readily methylated by adding the calcd. amt. of Me₂SO₄ and KOH, the product being insol in alkali, m. 170–1°. II, III, IV and V, form similar salts with the heavy metals as well as picrates. The picrates m. 234–5°, 210–1°, 191–2°, 198–9°, resp.

A. W. CONTRI

The influence of light, carbon monoxide and quinone on the methylene blue reduction. HIROSHI TAMURA, TATSUTARO HIDA AND KIYOSHI TANAKA. *Acta Phytochim.* 5, 119–55(1930)—AcCO₂H, "Kojii" acid, AcCl₂CO₂Et, phloroglucinol and resorcinol reduce methylene blue (I) in the absence of O₂ at a certain pH. Light energy is essential for this reaction. The presence of a mobile H atom in these substances (keto-enol tautomerism) is emphasized. The enzymatic reduction of I was studied with *Bacterium pasteurianum* and liver exts. It also was found to be speeded up by light. The I mol. is activated in this photochem reaction, not the H donor or the dehydrase.

The I reduction by chem. means or by liver exts. is not influenced by CO (II). However, I reduction by bacterial suspensions was found more or less prevented by II. The ratio I/II detcs the degree of inhibition. The light does not influence it. II apparently prevents the adsorption of I on the bacterial surface. AcOH dehydrase is not affected by II. Small amounts of quinone interfere with the chem. as well as the enzymatic reduction of I. Extensive experl. data are given and a special app. is described.

I. G. FRATER-FULNBOG

Thiodiazines. VI. PRAPULLA KUMAR BOSE AND BRENDRA KUMAR NANDI. *J. Indian Chem. Soc.* 7, 733-9 (1930), cf. *C. A.* 19, 831.—The mechanism proposed by Guha and Roy Choudhury (*C. A.* 23, 139), for the formation of the thiodiazine PhNH-NHC(N)N(CPh)CH₂S from the action of PhNHNHCSNHNH₂ on BrCH₂COPh is different from that suggested by R. who pointed out that varying quantities of thiazoles are formed in the above condensation (*C. A.* 20, 415). The formation of such thiazoles can be explained by the assumption of an intermediate of the type NH₂N=C(NH₂)SCCH₂COPh. Reactions between thiosemicarbazides and *p*-MeC₆H₄COCCH₂Br led, however, either to the formation of a thiodiazine or to a thiazole, but not to both as expected. 4-Substituted thiosemicarbazides gave thiodiazines whereas thiosemicarbazide gave a thiazole. One g. NH₂NHCSNH₂ and 2.3 g. *p*-MeC₆H₄COCCH₂Br (I) were shaken with 35 cc. abs. alc., and after refluxing for 20 min. the mixt. was cooled and added with Et₂O. The crude product was collected, washed with Et₂O, dried and dissolved in dil. HCl. After filtering and neutralizing with Na₂CO₃, the free base 2-keto-4-*p*-tolyl-2,3-dihydro-1,3-thiazole hydrazone (II) was pptd., which after recrystn. from MeOH, m. 178°; (HCl salt m. 133°, benzylidene deriv. m. 205° (identical with the condensation product of PhCH₂NHCSNH₂ and I), isopropylidene deriv. m. 230°; *Bz* deriv. m. 201°, *Ac* deriv. m. 195° (identical with the condensation product of AcNH₂NHCSNH₂ and I). The formation of the PhCH₂ and *Ac* derivs. rules out the formulation of II as a thiazole NH₂C(SCH₂C(C₆H₄)CH₃)NNH₂. II reacted with isothiocyanates,

forming addn. compds. of the type, NH₂C(SCH₂C(C₆H₄)CH₃)S₂C₆H₄NNHCSNH₂ (R = Me, Ph, *o*- and *p*-tolyl, m. 140°, 191°, 174° and 166° resp.), and with HCN to form HN(C(C₆H₄)CH₃)CH₂S₂C₆H₄NNHCSNH₂ (III), m. 218°. III dissolved in abs. alc., condensed with II to give (HN(C(C₆H₄)CH₃)CH₂S₂C₆H₄N)₂, m. 120°. NH₂NHCSNH₂ condensed with II to give 2-methylamino-5-*p*-tolyl-1,3,4-thiodiazine, MeNH-C₆H₄NH(C(C₆H₄)CH₃)CH₂S, m. 249° (HCl salt m. 195°), 2-isobutylamino homolog, m. 159°, was prepd. from II and iso-BuNHCSNH₂ (HCl salt m. 125°, *Ac* deriv. m. 167°), 2-phenylamino analog, m. 198° (HCl salt m. 135°, *Ac* deriv. m. 170°), 2-*o*-tolylamino compd., m. 179° (HCl salt, m. 171°, *Ac* deriv. m. 199°), and 2-*p*-tolylamino isomer, m. 187° (HCl salt, m. 199°, *Ac* deriv. m. 154°).

C. R. ADDINALL

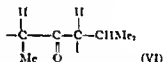
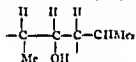
Correction to G. Hahn and W. Schuch Yohimbene alkaloids. VI. Two further secondary alkaloids of yohimbene. R. LILLIG. *Ber.* 63B, 2680 (1930)—H and S. state (*C. A.* 24, 5039) that L's α -yohimbene does not lose in wt. when dried at 105° and that the 15% loss cannot be ascribed to a const. hydrate or alcoholate. This is due to the use of aq. MeOH instead of EtOH for the recrystn., from 95% alc. the compd. seps with 1 mol. each of EtOH and H₂O. The loss in wt., compn. and titration results agree with the formula C₂₁H₃₁O₂N, EtOH·H₂O. The alc. can be detected by rubbing 2 g. of the base with 5 cc. H₂O disg. off 2 cc. and testing the distillate with 1 KI-NaOH or with BrCl-NaOH.

C. A. R.

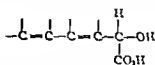
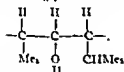
Sulfonation of alkaloids. ZENJIRO KITASATO AND KAKUJI GOTO. *Ber.* 63B, 2006-700 (1930)—In general, practically nothing is known as to the reactions which take place between alkaloids and concd. H₂SO₄. K. and G. have found that alkaloids having a HO (or MeO) group on the aromatic nucleus, with a free *p*-position, are sulfonated by concd. (not fuming) H₂SO₄ below 10°, the resulting acids generally form well-defined crystals and the SO₃H group always combines with the basic N to an inner salt. In this way were obtained *p*-aperine- (I), *apomorphine*- (II) and *sinomeninesulfonic acid* (III). These acids no longer give the diazo reaction and the SO₃H group is probably in the *p*-position to the phenolic HO group. Strychnine, which has no HO (or MeO) group, and brucine, in which the *p*-position to the HO group is not free, are not sulfonated under these conditions, and neither are those derivs. of sinomenine and thebaine, such as bromosinomenine, disinomenine and dithebaine (IV), in which,

judging from the failure to give the diazo reaction, the β position to the HO group is not free. I or I, the 2 position for the SO₃H group is assumed from analogy with nitropanaxinine. To test whether the color reaction of certain alkyls with H₂SO₄ is due to the formation of a colored compound or to a decoump. IV was studied first. It dissolves in concd. H₂SO₄ with deep red color, and when poured on ice forms a colorless soln. which with Na₂CO₃ regenerates the unchanged IV. I m. 101° (decomp.), easily sol. in Na₂CO₃, alkalis and H₂O; II, needles with 3H₂O m. above 300°. III, m. 265-7°. *Sinomenonesulfonate* and from *sinomenone hydrate* forms with 0.5 H₂O, m. 275-80°, gives in MeOH with NH₄OH HCl a *diastime*, m. above 200°, also obtained by sulfonation of *sinomenone hydrate diastime*. C. A. R.

Saponin of *Panax repens*, Maxim. IV. *Panaxsopogenin* S. Aoyama, J. Pharm. Soc. Japan 50, 1070-87 (1930), German abstr. 139-41, cf. C. A. 24, 125. A previously has reported on *panaxsopogenin* (I) and *panaxprosopogenin* (II) and shown that II is made up of *panaxsopogenin* (C₂₇H₄₄O₄) and glucuronic acid. Recently Kotake, Matsumura and Kimoto (J. Chem. Soc. Japan 51, 396-400 (1930)) have shown that saponin obtained from *Panax repens*, Maxim. is identical to that obtained from *Trapa chinensis*, L. var. *grabrescens*. K. also proposed to change the mol. formula of C₂₇H₄₄O₄ (Murray and Tanaka, C. A. 21, 3359) to C₂₇H₄₂O₄, but he could not account for the large mol. wt. of its Ac ester. A obtained slightly different results. Analysis of I indicates the mol. formula C₂₇H₄₄O₄, as previously reported. *Dihydroxopogenin* C₂₇H₄₄O₄·H₂O, m. 290-3°. I probably contains 2 double linkages, but no Me₂C=vinyl or vinylidene, ketone or aldehyde groups. I forms a *diphenyluratin* C₂₇H₄₄N₄, decoump. 131-40°, hence contains 2OH groups. Me ester of I, C₂₇H₄₄O₄, m. 198° contains 2 alc. OH and 1 MeO groups. *Monooctylsopogenin* (III) C₂₇H₄₄O₄, m. 262-45°, and corresponds to the C₂₇H₄₄O₄ of K. Only 1OH group is acetylated, the other is probably tertiary alc. *Diethylsopogenin anhydride* (IV) m. 315-8°. IV, when treated with KOH followed by HCl, gave III. Me ester of III, C₂₇H₄₄O₄ (V), m. 221°. Oxidation of V with CrO₃ gave a diketone, C₂₇H₄₄O₄ (VI), m. 192-3°, and an acid decoump. 215-8°. This acid is under further investigation. VI shows neither the Liebermann nor the C₂₇H₄₄O₄ reaction, hence it contains no —COMe group. VI is an unsatd. ketone. Reduction of VI by Clemmensen's method gave the original compd. *Diastime* of VI, C₂₇H₄₄O₄·N₂, decoump. 237°, probably contains a —C(CO)C(CO)C(CO)C(CO)— linkage. VI forms a *monosemicarbazone*, C₂₇H₄₄O₄·N₂, which does not m. 200°. V gives neither the diazone, pyrrole or pyrazoline reactions, hence the position of the C=O group must be either at 1 and 5 or further apart. A suggests the following structure to explain the position of OH and CO groups in I and VI:



Sapon. of V gave the Me ester, m. 195°, while I gave a compd., C₂₇H₄₄O₄ (VII), m. 203°, which did not depress the m. p. of I. VII is probably an isomer of I and is called α -*isopanaxsopogenin*. III and VI also give VII. The exact mechanism of the formation of the isomer is not known. I is stable toward concd. HCl, even on heating, and also alc. HCl. When treated with dry HCl in excess MeOH I gives a compd. m. 268° (VIII). VIII is probably another isomer of I and is called β -*isopanaxsopogenin*. A suggests the following partial structure for I:



F. I. NAKAMURA

Digitalis glucosides. III. Glucosides of *Digitalis lanata*, Sydney Smith, J. Chem. Soc. 1931, 23-5, cf. C. A. 24, 2753, 25, 704.—In acid to digoxin, which was previously reported, gitorin, m. 235° (cor. decomp.), [α]_D²⁰ 3.5° (C₂₇H₄₄N, c 1.02), has been isolated from the leaves of *Digitalis lanata*, this has hitherto been found only in the leaves of *D. purpurea*. Other glucosides in the leaves are being examd.

C. J. WEST

Brazilin and hematoxylin questions. XI. Hydroxybenzylchromanones, PAUL PEPPER, ERWIN BREITH AND HANS HOYER, J. prakt. Chem. 129, 31-51 (1931); cf.

C. A. 24, 4301—Chromanone (I) (0.5 g) and 0.46 g p -MeOC₆H₄CHO in abs EtOH with dry HCl at 0° give 0.51 g 4'-methoxybenzalchromanone, yellow, m 131°, 0.0001 mol in 1 cc concd H₂SO₄ gives a blood red color. I and 0.55 g vanillin Me ether give 0.28 g 3',4'-di-MeO deriv, pale yellow, m 117°, color in H₂SO₄ blood red. I and 0.5 g piperonal give 0.45 g 3',4'-methylenedioxy deriv, yellow, m, 134.5-7°, color, Bordeaux red. I and 0.5 g vanillin give 0.22 g 3'-methoxy-4'-hydroxy deriv, yellow, m 126-9°, color, deep blood red, Ac deriv, golden yellow, m 160° (84% yield). I and 0.47 g 3,4 (110), C₆H₄CHO give 0.42 g of 3',4'-di-HO deriv, golden yellow, m 224-5°, color, wine red, di Ac deriv, m 166° (52% yield). 7-Methoxychromanone (5 g) and 4.3 g isovanillin in abs EtOH with HCl give 9.2 g 4',7-dimethoxy-3' hydroxybenzalchromanone, light yellow, m 153-4°, color, deep orange, Ac deriv, pale yellow, m 139-40°; catalytic reduction of the Ac deriv gives the benzyl deriv, red, m 123-4°, oxime, m 160-1°, Ac deriv, m 90-1°. 7-Hydroxychromanone and vanillin Me ether give 3',4'-dimethoxy 7 hydroxybenzalchromanone, brown, m 245-9°, color, dark red, Ac deriv, yellow, m 151-2° 3',4'-Methylenedioxy deriv, brown, m 234-6°, color, blood red, Ac deriv, yellow, m 132.5-4° 3'-Methoxy-4',7-dihydroxy deriv, golden yellow, m 230-1.5°, color, dark red, Ac deriv, m 151-2.5° 3',4',7-Trihydroxy deriv, yellow, m 250-3°, color, blood red, tri Ac deriv, pale yellow, m 132-4°. The corresponding benzyl deriv, yellow, m 201-2°, color, light wine yellow, Ac deriv, m 117°, 7,8-Dimethoxychromanone with AlCl₃ or concd HBr give the 7,8-di-HO deriv, crystg with 1 mol H₂O, m 183-8.5°, color, citron-yellow, di Ac deriv, m 111°, semicarbazone, pale brown, m 224-5° (decomp). With p -MeOC₆H₄CHO and HCl in EtOH there results 4'-methoxy 7,8-dihydroxybenzalchromanone, greenish yellow, m 192°, color, deep brownish red, Ac deriv, light yellow, m 140°. 3',4'-Di MeO deriv, crystg with 1 mol H₂O, light yellow, m 174-5°, color, carmine red, Ac deriv, pure yellow, m 183°, the corresponding benzyl deriv, reddish needles m 177°, Bz deriv, m 177.5-8°. 3'-Methoxy-4',7,8 trihydroxybenzalchromanone, crystg with 1 mol H₂O, yellow, m 200-7°, attempts to remove the H₂O by heating at 17 mm over P₂O₅ caused decomp, Ac deriv, m 180-1° 3',4',7,8 Tetra-HO deriv, light yellow, begins to decomp at 230° and is complete at 265°, color, brick red, Ac deriv, m 166-7°; the corresponding benzyl deriv, yellow red, decomp above 200°, Ac deriv, m 107.8°

C. J. WEST

Santonin. KARL JOSEPHSON *Stensk Farm Tids* 35, 29-33(1931)—The lactone ring of santonin is opened by NH₃ in alc giving santoninic acid amide. This reaction is easily reversed. J gives a general discussion on the constitution of santonin with bibliography up to Clemo and Haworth (C. A. 24, 4040).

A. R. ROSE

Hydrolysis of benzoylated amino acids and polypeptides. STEFAN GOLDSCHMIDT AND WILHELM FEYER *Ann* 483, 190-216(1930)—The hydrolysis of various benzoylated mono- and diamino acids, aminohydroxy acids and polypeptides with NaOH (2% usually) at 25° and 100° and H₂SO₄ at 100° is reported. The rate of alk hydrolysis of the monoamino acids increases with rise in temp and HO ion concn, for a definite HO ion concn, the coeff decreases with increasing length of the C chain. Alk hydrolysis of hippuric acid is increased at both 25° and 100° by the addn of NaCl but is decreased by added glycine. The O Bz group is eliminated from O,N dibenzoylserine and tyrosine at 25°, at 100° both Bz groups are readily eliminated from the former compd. Alk hydrolysis of dibenzoylarginine causes the very ready elimination of 1 Bz group with the production of α benzoylornithine, the velocity of hydrolysis of the second Bz group is measurable when 8% NaOH is used. Similarly, α -dibenzoyllysine (I) is first hydrolyzed to the α Bz deriv. With benzoylated glycyl, diglycyl and alanylglycylglycines the peptide linkages are rapidly broken, during this hydrolysis little BzOH is produced. The rate of hydrolysis of benzoylated monoamino acids by 70% H₂SO₄ at 100° increases with the length of the C chain. Dibenzoylarginine behaves similarly with acidic hydrolytic agents as with alk, but the behavior of I is reversed the α Bz group being eliminated first. The peptide linkages are first broken when the above benzoylated peptides are hydrolyzed with 70% H₂SO₄. The velocity of hydrolysis of benzoylproteins is greater with 30% than with more concd H₂SO₄, the rate corresponds with the sp cond of the acid. The velocity of hydrolysis of hippuric acid by 70% H₂SO₄ is increased by the addn of glycine.

C. J. WEST

A higher alcohol obtained from "Cortinellus Shitake." MIZUHO SUMI *Bull Inst Phys Chem Research (Tokyo)* 9, 966-9(1930), Abstracts 92 (in English) published with *Sci Papers Inst Phys Chem Research (Tokyo)* 15, Nov 278-81—A higher solid alc C₁₅H₃₂O is obtained from finely powd *Cortinellus Shitake*, needles or prisms, sol in Et₂O, C₆H₆, CHCl₃, not easily sol in EtOH, Me₂CO, etc., m 245-55°, $[\alpha]_D^{25}$

—22° Acetate, m 260-5°, $[\alpha]_D^{25}$ -30.2° Benzoate, m 210°. Palmitate, m 171-2°. The spectrum shows no special absorption band. K. KONDA

The bile acids. XXXI. The constitution of pseudocholoidanic acid. HEINRICH WIELAND, LUDWIG FRITZ AND ELIZABETH DANE *Z physiol Chem* 104, 107-14 (1931).—The study of pseudocholoidanic acid has been resumed with the hope of determining the location of 2 remaining C atoms in the bile acids. The tetrabasic pseudocholoidanic acid $C_{21}H_{34}O_8$ is formed along with the pentabasic choloidanic acid, $C_{21}H_{32}O_8$, when desoxybilanic acid is oxidized by HNO_3 . The former view that this oxidation breaks open ring III is now abandoned in view of the subsequent location of a Me group at C_{11} . Pseudocholoidanic acid forms a neutral tetra Me ester, but no keto groups can be demonstrated, hence the remaining 20 atoms must be present in ether or lactone linkage. When heated it loses CO_2 and $2H_2O$ and forms a pyropseudocholoidanic acid $C_{21}H_{30}O_7$ which is monobasic and yields a mono-Me ester, m 102°. Treatment of the pyro acid with alc KOH opens up a lactone ring with formation of a dibasic acid, $C_{21}H_{32}O_8$, m 212°, from which a di-Me ester was obtained with CH_3N_3 . The mono-Me ester of pyropseudocholoidanic acid was also obtained by heating the di-Me ester of pseudocholoidanic acid. A by product in the prepn of the pyro acid was a dibasic acid, $C_{21}H_{32}O_8$. The open pyro acid is tribasic when titrated hot, due to hydrolysis of an anhydride grouping. Desoxybilanic acid in AcOH takes up 2Br, yielding 50-60% of dibromodesoxybilanic acid m 215° (decompn). Treatment of this with $N KOH$ in MeOH removes HBr with formation of bromodesoxybilanic acid, m 218° (decompn), in 80% yield, tri-Me ester, m 93°. Both of these brominated acids revert to desoxybilanic acid when reduced with Zn and HCl. Oxidation of the unsatd bromo acid with HNO_3 yields pseudocholoidanic acid. This new manner of obtaining pseudocholoidanic acid proves it to be a lactonetetracarboxylic acid, in which ring II has been opened by breaking the linkage between C_6 and C_7 , since the oxidation product contains no Br. Formulas representing the characteristic groupings of pseudocholoidanic acid and its pyro acid are proposed. XXXII. Attempts at a new breakdown of desoxybilanic acid. ELIZABETH DANE AND HEINRICH WIELAND *Ibid* 119-23.—In dibromodesoxybilanic acid (vide supra) both Br atoms are on C_6 . One is easily removed as HBr, but the 2nd is tenaciously held as would be expected from the presence of an adjacent double linkage. Oxidation of the unsatd bromodesoxybilanic acid by alk $KMnO_4$ converts it into the tribasic 5-hydroxy-5-ketodesoxybilanic acid, $C_{21}H_{34}O_8$, m 233-5°, orime, m 200°. Esterification with CH_3N_3 yields a tri-Me ester, m 154°, while esterification with MeOH-HCl yields an isomeric tri-Me ester, m 178°. The difference between these isomers is not understood. The HIO acid when boiled with AcOH and HCl loses H_2O and forms a diketolactonedicarboxylic acid, $C_{21}H_{34}O_8$, m 258-60°. Catalytic hydrogenation of the latter yields 70% of a 6-ketodesoxybilanic acid, $C_{21}H_{34}O_8$, m 255°, tri-Me ester, m 108°. Oxidation of the above HIO acid with fuming HNO_3 gave a pentabasic acid, $C_{21}H_{32}O_8$, m 180°. A. W. DOX

Taurolithocholic acid from chicken bile. TADA0 HOSIZIMA, HIROSI TAKATA, ZIRO URAKI AND SIGETOSI SANOYA *J Biochem (Japan)* 12, 393-7 (1930).—Chicken bile is freed from mucus by alc and is concd on the water bath to the original vol. of 1 l. $FeCl_3$ is then added until the acid soln gives no further ppt. The ppt. is filtered off and is decompd with warm Na_2CO_3 . The filtrate is evapd and the dry residue extd with alc. The alc soln is decolorized with animal charcoal and evapd, the residue again taken up in alc and decompd with H_2SO_4 and mixed with ether. The flocculent ppt is now filtered off. This ppt is dissolved in dil Na_2CO_3 , gradually acidified with HCl and extd by shaking with ether. The ppt thus formed is recrystd several times from dil alc and gives a yield of 2.13 g. These crystals are not sol in H_2O , ether, petroleum ether or benzene, but do dissolve easily in alc, AcOH or Me_2CO , they give the Pettenkofer reaction and m 210-1°. In abs. alc soln (0.54%) they have a sp rotation of 65.74°. Its elementary compn corresponds to the formula $C_{26}H_{40}OCONHCH_2CH_2SO_3H$. Heated for 24 hrs under a reflux with 5% NaOH, it gives a mass of beautiful roset shaped needles. The Na salt was converted to a Ba salt which was recrystd from dil alc by a special procedure. Analyses on this Ba salt show that the bile acid is related to lithocholic acid, and is identified as tauroisolithocholic acid. S. MORGULIS

Magnetic rotatory power of hydrocarbons in the gaseous state (MALLEMANN GABIANO) 2. Trichloromethyl perchlorate (BIRCKENBACH, GOUBEAU) 6. Recent experiments on the pyrolysis of CH_4 (WHEELER, WORD) 21. Kinetics of the formation of malonamide from ethyl malonate and NH_3 in homogeneous solution (BAILEY) 2. Formation of CH_4 during the electrolysis of K acetate and the mechanism of Kolbe's elec-

trosynthesis (SHUKLA, WALKER) 4. Natural and synthetic rubber (MIDGLEY, HENNA) 30. Apparatus for effecting gas reactions such as catalytic production of MeOH (U S pat 1,700,853) 18.

FILIN, CARLETON: Hydrogenation of Organic Substances 3rd ed. New York: D Van Nostrand Co, Inc 940 pp M. 15 Reviewed in *Ind Eng Chem* 23, 342 (1931).

Hydrocarbons. ALFONS DESCHAUER Ger 515,177, Dec. 1, 1926 Partly hydrogenated polycyclic hydrocarbons are prepd by treating the acid sulfates of cyclic hydrogenated alcs with aromatic hydrocarbons in the presence of H_2SO_4 . Thus, cyclohexanyl sulfuric acid and CaH_2 yield phenylcyclohexane. The acid sulfates used as initial materials need not be separately prepd. Thus, an aromatic hydrocarbon may react, in the presence of H_2SO_4 , with a cyclic unsatd hydrocarbon such as cyclohexene, or with a cyclic hydrogenated alc. Examples are given.

Hydrocarbons. I. G. FARBERINO A-G Fr. 603,054, Mar 31, 1930 Hydrocarbons of high b. p. are converted into hydrocarbons of lower b. p. by cracking or hydrogenation in the presence of catalysts composed of metal compds of the type of acetylacetonates, with or without other catalysts.

Partial combustion of hydrocarbons. JOHANN S. OHLENDIECK-DOLGE Ger 514,394, May 4, 1928 This relates to the manuf of mixts contg CO and H by treating hydrocarbons with air or other oxidizing gas at a raised temp in the presence or absence of catalysts. The improvement consists in using an alloy of Fe, Cr, and Ni, that has been fused *in vacuo*, for constructing the parts of the reaction chamber that come into contact with the gases.

Condensing diolefins with other hydrocarbons. I. G. FARBERINO A-G Brit. 335,512, July 5, 1929 In the further application of the general type of reactions described in Brit 315,312 (C A 24, 1649), isoprene, toluene and Na are heated in an autoclave at 160° (yielding 3-methyl 5 phenyl-2-pentene or 2 methyl-5-phenyl 2-pentene β,γ -Dimethylbutadiene, toluene and Na yield 2,3-dimethyl-5-phenyl 2-pentene, and butadiene, tetrahydronaphthalene and Na yield butenyltetrahydronaphthalene Cf C A 24, 3247.

Aliphatic compounds. HENRY DERYFUS Fr 602,736, Mar 25, 1930 Org O compds are synthesized by passing the reaction gases through a no. of reaction vessels in series in an open circuit. An example is given of the prepn of MeOH from water gas Cf C A 25, 963.

Organometallic compounds. LOUIS P. V. LECOG Fr 602,949, July 1, 1929 Salts and complex metallic compds of dithiocamphocarboxylic acid which are sol in org solvents, particularly in animal and vegetable oils, are prepd by double decompn or by the action of the acid on oxides. Examples are given of the prepn of the Cu, Au, Ag and Hg salts and of a complex compd contg As which is sol in olive oil, by the reaction of dithiocamphocarboxylic acid with β -hydroxy-*m* aminophenylarsonic acid Cf C A 24, 2138.

Diazo compounds. I. G. FARBERINO A-G (Erich Lehmann, inventor) Ger 515,205, Nov 6, 1928. Diazo compds of unsulfonated diphenylamine derivs. are salted out from their solns at atm temp by means of neutral, unhydrolyzed salts in the presence of a weak acid, e. g., H_2BO_3 , and a compd capable of reacting with HNO_2 , e. g., urea. The oxidation products which sep after the diazotization should be removed before salting out. The products are stable in the cold and in warm, damp air. Ex. amples are given.

Alcohols from olefins. N-V. DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ Brit 335,551, May 23, 1929 Olefins (including those contg 2 or more double bonds) are hydrated in the vapor phase in the presence of a metal of the Pt group, Au, Ag, Cu, Fe, Ni, Co, Cr, Ta, V, W, Mo or Mn or their salts or other compds (most suitably compds which during the reaction yield the metals such as oxalates, carbonyls and the like, although compds such as Bi vanadate, Cu vanadate or Cu phosphate also may be used). Several examples with details of procedure are given for treating ethylene and propylene.

Cyclic aldehydes. I. G. FARBERINO A-G (Georg Kalischer, Heinz Scheyer and Karl Keller, inventors) Ger 514,415, Feb 3, 1927. The aldehyde group is introduced into oxygenated cyclic compds by treating these with formyl derivs of secondary amines and a chloride or oxychloride of P or S. The reaction may be effected in an inert solvent. Thus, β -naphthol may be added to a benzene soln of the product obtained by treating formylmethylaniline with $POCl_3$ while cooling. 2 naphthol 1-

aldehyde is obtained from the mist after standing for about 20 hrs. Other examples are given also. The method may be used to prep aldehyde deriva. of anthrone, benzanthrone, naphthostyryl, hydroxanthracene and hke compds. Cf. C. A. 24, 3500.

Polycyclic ketones. I. G. FARRINGTON, A. G. (Otto Nickschum and Walter Bernth, inventors). Ger. 514,110, Dec. 10, 1925. Polycyclic ketones are prepd from aromatic or heterocyclic ketones having halogen in the α or β position to the C=O group by treating them at a temp. of temp. with alkali or alk. earth salts or with leekly alk. aryl haliding agents. The ketones may be treated in the gas or liquid phase and in the presence or absence of catalysts, e. g. Cu or CrCl_3 . Thus fluorenone may be prepd by heating α -chlorofluorenone with NaClO_2 at 400° for some hrs. Other examples are given also. Cf. C. A. 1, 25, 1260.

Acetals. CONSORTIUM FOR VERFAHRUNG, INO. G. M. B. H. (Hans Deutsch and Willy O. Herrmann, inventors). Ger. 513,700, Jan. 20, 1926. Adbn. to 502,111. The method of Ger. 503,411 et. 1, 21, 1700 is modified by using abs. contg. water. The water is converted into H_2 which may be fractionated out of the reaction products or removed by washing, e. g. with the alc. used. The yield of acetals may be increased by partly condensing the reaction products and returning the condensate to the reactn. vessel or by circulating the products through the vessel. Examples are given.

Aldehyde bases. PROMETHEE ROCHER, S. A. ANON. Belg. 578,810, July 11, 1916. Hydroaromatic aldehydes are condensed with $\text{C}_6\text{H}_5\text{OH}$ and aliphatic amines.

Aldehyde amines. THE NATURALIA COMPANY, CO. U. S. 1,912,835, Mar. 27, 1930. The determination of condensation products of aldehydes with amines such as heptyl aldehyde with PhNH_2 is retarded by treating the products with a polyhydric alcohol or polyhydric substituted aromatic compd. such as catechol, resorcinol, pyrogallol, dinitro- α -naphthol, hydroquinone or its mono- or di-ethers, etc.

Esters. OTTAVIANI (to Deutsche Gold- und Silber-Scheideanstalt vorm. Rossmeyer). U. S. 1,791,218, Feb. 3. In producing esters such as MeCOAc , aliphatic acids such as HOAc and alcs. such as MeOH are caused to react together at temps. of about $100-180^\circ$, corresponding to working pressures of 2-10 atm. In the presence of catalysts such as H_2SO_4 or NaOH , which are capable of producing moderate H ion concn., and the ester formed is continuously removed from the reaction mist, as produced.

Vinyl esters. WILLY O. HERRMANN, HANS HERZFELD and EUGEN HAHN (to Consortiun für Elektrochemische Industrie). U. S. 1,790,920, Feb. 3. An excess of C_2H_4 is passed through a monoelectrolyte aliphatic acid such as HOAc in the presence of not more than about 1% of a Hg compd. such as Hg acetate and the resulting vinyl ester is removed from the reaction mist by entrainment with the escaping excess of C_2H_4 , and is recovered by rectification. Several examples with details of procedure are given. Cf. C. A. 24, 1527.

Conversion products of vinyl esters. CONSORTIUM FOR VERFAHRUNG, INO. G. M. B. H. (Hans Deutsch and Willy O. Herrmann, inventors). Ger. 515,107, Mar. 11, 1927. Ketone and other ketones, AcH , acid anhydrides and vinyl ethers are among the products obtained by passing vaporized vinyl esters at a temp. between 250° and 500° and at atm. or raised pressure through a packed reaction vessel which may contain a catalyst. Thus, the reaction vessel may be packed with wood charcoal on which ZnO has been deposited. Examples are given.

Polyvinyl alcohol esters. I. G. FARRINGTON, A. G. U. S. 1,902,810, Mar. 27, 1930. Sol. mixed esters of polyvinyl alc. are made by the reaction simultaneously or successively of the alc. mists with acid halides or condensation products obtained by elimination of halogen hydride from halide acids. The products are converted to an insol. form by heat. Thus a mist of polyvinyl alc., HCl , naphthalene acid chloride, PhCl and pyridine bases are heated to $35-40^\circ$ for 24 hrs. The product is sep'd by MeOH and filtering. It can be formed into a film from its soln. in acetone. Other examples are given.

Alkyl chlorides. F. I. DUFOUR DE NEMOURS & CO. U. S. 1,902,790, Mar. 20, 1930. Alkyl chlorides are made by bringing a mist of vapors contg. alc. and HCl into contact with an aq. soln. of a metallic chloride such as ZnCl_2 or FeCl_3 at a temp. of about $100-160^\circ$. The mist of vapors may be formed by the reaction of NaCl , H_2SO_4 and alc. at a high temp. An example is given of the prepn. of EtCl . Cf. C. A. 25, 3001.

Calcium benzoate. THE GORGONICUM A. G. U. S. 1,902,821, Mar. 21, 1930. Phthalic anhydride is ground with an aq. suspension of Ca(OH)_2 and heated to a temp. above 350° . The mist is introduced gradually into the reaction chamber and a dilg. agent such as some of the material already transformed may be added to the mist.

Sulfonates of pyridine and its homologs. I. G. FARRINGTON, A. G. Belg. 575,817, Nov. 20, 1924. In the sulfonation of pyridine and its homologs (suitably with fuming

H_2SO_4 , Hg or a Hg salt such as the sulfate is used as catalyst. Examples are given of the sulfonation of pyridine and of α -picoline. Previous use of various other catalysts also is mentioned.

Stabilizing solutions of salts of aromatic stilbinic acids. I G FARBERNIND A-G (Hans Schmidt, inventor) Ger 515,113, Feb 25, 1927. A polyhydric alc., e.g., glycerol or glucose, is included in the solns. The alc and the salt may be present in equal amounts.

Amine derivatives. IMPERIAL CHEMICAL INDUSTRIES, LTD Fr 693,180, April 2, 1930. N,N' -Thio derivs of primary and secondary amines are made by the interaction of a primary or secondary amine with a chloride of S in aq, neutral or alk suspension. Examples are given of the prepn of $\text{Li}_2\text{NSSNEt}_2$ from Et_2NH and $(\text{CH}_3)_3\text{NSSLi}$ from piperidine.

Anthraquinone derivatives. I G FARBERNIND A-G (Max Kugel, inventor) Ger. 513,025, June 12, 1927. Mixts of α - and β -aroylaminoanthraquinones or their halogen derivs are sepd by treatment in alc. suspension with caustic alkali at a temp sufficiently low to avoid sapon. The α -derivs remain undissolved, while the β derivs. go into soln., and can be recovered by acidification. Examples are given. Cf. C. A. 25, 965.

Benzanthrone derivatives. I. G FARBERNIND A-G (Heinrich Neresheimer and Willy Eichholz, inventors) Ger 515,327, Aug 11, 1929. Addn to 431,774. According to Ger 431,774, Bz -1, Bz -1'-Dibenzanthronyls are prepd by treating benzanthrone, or its derivs having the 2 position free, with acid oxidizing agents. It is now found that the reaction proceeds differently if oxidation is effected with MnO_2 and moderately dil H_2SO_4 , products sol in alkalis and in cold bisulfite soln being obtained. Like products can be prepd from benzanthrone derivs. substituted in the 2 position, and in treating these, other acid oxidizing agents may be used. Examples are given. Cf. C. A. 25, 1262.

Naphthocarbazole derivatives. I G FARBERNIND A-G (Albert Schmelzer, Fritz Ballauf and Heinrich Hefner, inventors) Ger 514,420, Feb 9, 1929. See Fr. 689,515 (C. A. 25, 906).

Alkyl cyclohexanols and cyclohexanones. SCHIERING-KAHLBAUM A-G (Walter Schoeller and Hans Jordan, inventors) Ger 515,112, Jan 27, 1928. Addn to 512,831. The hydrogenation process described in Ger 512,831 (C. A. 25, 1259) is carried out in the presence of a surface catalyst, e.g., Al-Mg hydrosulfate. The hydrogenation catalyst (Ni) may be deposited on the surface catalyst. Examples are given. Cf. Ger 512,235 (C. A. 25, 974).

Phenols. I G FARBERNIND A-G (Friedrich W. Stauf and Gerhard Kröner, inventors) Ger 513,798, Mar 20, 1928. See Brit 308,220 (C. A. 24, 130).

Improving the odor of phenols. GEWERKSCHAFT MATTHIAS STINNES, Ger 515,407, Oct 29, 1927. Addn to 514,341. The method of Ger 514,341 (C. A. 25, 1336) is modified by partially etherifying the phenols instead of partially esterifying.

Indoles. IMPERIAL CHEMICAL INDUSTRIES, LTD Fr 692,590, Mar. 24, 1930. See Brit. 330,332 (C. A. 24, 5770).

Styrenes. THE NAUGATUCK CHEMICAL CO Fr 693,876, April 14, 1930. Styrenes are prepd by heating an alkylbenzene and C_6H_6 . A substance such as S or a mixt of CuI , Cl_2 and S capable of forming a stable compd. with the H liberated may be added.

Acetic acid. HENRY DREYFUS Fr 692,735, Mar. 25, 1930. The vapor of CH_3O alone or mixed with CO is submitted to the action of heat under a pressure of 100-300 kg or more. A temp. of 300-400° is preferred and the CH_3O may be used in the form of formal and in the presence of steam. Catalysts such as inorg acids may be used.

Aliphatic anhydrides. HENRY DREYFUS Fr 692,537, Mar 21, 1930. Aliphatic anhydrides, particularly Ac_2O , are made by passing the vapor of aliphatic acids, particularly AcOH , in contact with a heated mass of Cu or other metal of good heat cond. Catalysts known to favor the reaction may be mixed with the Cu.

Aliphatic anhydrides. HENRY DREYFUS Fr 693,483, April 7, 1930. Aliphatic anhydrides, particularly Ac_2O , are prepd by the thermal decompn. of aliphatic acids, particularly AcOH , in the presence of a catalyst composed of or contg. one of more anhydrides or acids of As or Sb or the salts of such acids, such as arseniates or antimonates of Na, K or Ca. The preferred temp is 350-700°. Cf. C. A. 25, 971.

Ethyl alcohol and other oxygenated organic compounds. H. DREYFUS Brit 335,631, July 2, 1929. EtOH and other products such as higher alcs., esters, aldehydes and ketones are prepd by heating MeOH vapor in the presence of catalysts comprising compds. of alkali or alk. earth metals in which there are the sole metallic compo-

nents (excluding Mg compds) such as oxides, hydroxides and alcoholates of Na, K, Ba and Ca, preferably at 400–600°, although higher or somewhat lower temps. may be employed. Various details of app and procedure are described.

Methyl and ethyl acetates. DEUTSCHE GOLD UND SILBER SCHMIDDEANSTALT VON ROESSLER (Otto Fuchs, inventor) Ger 513,943, Jan 16, 1927 and 514,744, Mar 22, 1927. Adds to 507,205. In applying the method of Ger 507,205 (C A. 25, 524) to the esterification of crude pyroligneous acid the acid is preliminarily purified by heating under pressure (513,943). This purification treatment may be effected in the presence of salts, e.g., 20–40% of a bisulfate (514,744).

Benzanthrone, etc. IVAN GUBELMANN, ROBERT J. GOODRICH and EDWARD T. HOWELL (to Newport Chemical Corp.) U. S. 1,791,309, Feb 3. In producing anthraquinone derivs. such as benzanthrone or methyl benzanthrone, a starting material such as anthraquinone or methyl anthraquinone is caused to react in concd H_2SO_4 with a Cu salt such as $CuSO_4$ and a metal such as Zn or Al which is higher than Cu in the electromotive series (suitably at 25–30°).

3-Nitro-5-ketotetrahydronaphthalene. I. G. FARBENIND A. G. (Fmil Laage, inventor) Ger 515,111, Apr 12, 1929. The anil of 5-ketotetrahydronaphthalene is nitrated and the product decompd with dil mineral acid. An example is given.

α -Naphthol. I. G. FARBENIND A. G. Fr 694,596, April 9, 1930. α -Naphthol is prepd by heating a ketotetrahydronaphthalene with S at about 200°. A catalyst such as Se, finely divided Sn, ZnO, arsenious anhydride or diphenylguanidine may be used.

1,4,1',4'-Tetrahydroxy-2,2'-dianthraquinonyl. I. G. FARBENIND A. G. (Robert J. Schmidt, Berthold Stein and Kurt Damberger, inventors) Ger 515,114, Sept 22, 1929. Quinizarin is treated with piperidine or its homologs or analogs at atm. or slightly raised temp. An example is given.

1,4,4'-Trihydroxy-2,2'-dianthraquinonyl-3,1'-oxide. I. G. FARBENIND A. G. (Robert J. Schmidt, Berthold Stein and Kurt Damberger, inventors) Ger 515,115, Oct. 2, 1929. 1,4,1',4'-Tetrahydroxy-2,2'-dianthraquinonyl in nitronaphthalene or other nitro compd of high b.p. is boiled for a short time. An example is given. Cf preceding abst.

Triazine derivatives. I. G. FARBENIND A. G. Brit 335,783, Oct. 31, 1928. See Ger. 501,087 (C A. 24, 4530).

Mercaptobenzothiazole, etc. IMPERIAL CHEMICAL INDUSTRIES, LTD., and K. H. SANDERS. Brit. 335,567, March 26, 1929. Mercaptobenzothiazole is prepd by introducing a soln of S in CS_2 mixed with $PhNH_2$ into a reaction vessel heated to 200–275° and under pressure, and allowing the mixt., on completion of the reaction, to overflow into a second chamber when it is drawn off. App. is described adapted also for other reactions of liquids and overflow of the reaction products.

β -Hydroxy- α -picoline. I. G. FARBENIND A. G. Brit 335,818, Nov. 26, 1928. α -Picoline- β sulfonic acid (which may be prepd as described in Brit. 335,817) is fused with an alkali such as KOH. Production of an azo dye by coupling it with diazotized p -nitroaniline in Na_2CO_3 soln. is described.

Polymerized vinyl chloride. IWAN OSTROMISLENSKY (to L. A. Van Dyk) U. S. 1,791,009, Feb 3. For obtaining the β -modification of polymerized vinyl chloride, the vinyl chloride is polymerized until the δ modification is formed, followed by addn of aniline, pptn with alc. and dissolving in C_2H_5Cl .

Purpurin. HENRY R. LEE and IVAN GUBELMANN (to Newport Chemical Corp.). U. S. 1,790,932, Feb 3. A soln of 3-chloroquinizarin in concd H_2SO_4 is heated at about 220° for a prolonged time (suitably about 20 hrs.).

Quinizarin. IVAN GUBELMANN and HENRY J. WEILAND (to Newport Chemical Corp.). U. S. 1,790,915, Feb 3. In prepg a quinizarin, a p -chloroaniline is diazotized in H_2SO_4 , the diazo compd thus formed, without isolation, is reacted upon in substantially anhyd. H_2SO_4 with phthalic anhydride in the presence of H_2BO_3 at a temp. of about 200° and this temp. is maintained until condensation is substantially complete, followed by cooling and diln. with water to ppt. the quinizarin compd.

Leuco indigo. IMPERIAL CHEMICAL INDUSTRIES, LTD. Fr. 693,469, April 7, 1930. Leuco indigo is made in a continuous manner by passing a paste of alk. indigo over a Ni catalyst on a rigid or semi rigid support composed of a granular reticular or tangled mass, in the presence of H under pressure. A temp. of 120–140° is used and the catalyst may be activated with Al_2O_3 .

11—BIOLOGICAL CHEMISTRY

PAUL R. HOWE

A—GENERAL

FRANK P. UNDERHILL

Chemical properties of blood pigments and related substances. A. STREEMAN. *Chem. Weekblad* 27, 170-4(1930)—The general chem. and optical properties of hemoglobin and its derivs. are described, and the variations exhibited by various types of blood are discussed. B. C. A.

Phiothione as hydrogenase. J. DE REY-PAILLIARD. *Compt. rend. soc. biol.* 99, 1700-2(1929)—Egg white, dild. with an equal vol. of water, was used as a source of phiothione, the addn of S liberates H_2S in 30 min. The phiothione formed contains the -SS- linking characteristic of cystine. The reaction does not occur after heating with dil. acid or alk. media; further, the ppt. formed by heating does not give H_2S after the addn of S. It is maintained that phiothione is an enzyme. B. C. A.

The Spierer lens and what it reveals in cellulose and protoplasm. WILLIAM SEIFRITZ. *J. Phys. Chem.* 35, 118-29(1931)—The Spierer lens is an oil immersion objective in which is inserted a mirror, smaller than the lens, which reflects all direct light coming into the objective. A diaphragm controls the admission of light. Scattered light is picked up by the lens around the mirror. The structure of cellulose and protoplasm as revealed by the Spierer lens is that of tiny rods or supermicelles arranged end to end to form long parallel striae. F. E. BAOWN

The chylomicron emulsion. S. DEW. LUDLUM, A. E. TAFT AND R. L. NUGENT. *J. Phys. Chem.* 35, 269-88(1931)—Chylomicrons are oil droplets in the blood. Their diams. are 1μ and smaller. That they are surrounded by a protective protein layer is indicated by: (a) their max. flocculation at pH 4.8-5.2, (b) the reversal of their charges at pH 4.8-5 and (c) coalescence to form larger droplets at a point in acid or alk. concn. at which protein would be salted out. Mixed serum albumin and globulin would behave in exactly this way when acting as protective layers on colloidal particles. The application of the Mudd interface technique gives results similar to those observed for fat globules in milk. These milk globules are believed to be surrounded by protein. The question of the assimilation and storage of fats and the behavior of lipid droplets in clinical tests, as those for syphilis, make this an important subject for study. The known behavior in these fields is in accord with the assumption of a protein sheath. F. E. BAOWN

The end point of tryptic action. J. TEMMINCK GROLL. *Pharm. Weekblad* 68, 109-21(1931)—A study is made of the factors which might be expected to influence the extent of hydrolysis of casein and gelatin by trypsin preps from hog pancreas. Ordinarily about 20% of the total N is converted in 1 week into amino N which can be titrated with CH_3CO . After several weeks the conversion may be as high as 30%, but this further increase does not represent enzymic action since the control without enzyme behaves similarly. The enzyme does not become inactivated since it readily attacks fresh addns of substrate. Moreover, the addn. of fresh enzyme does not increase the extent of hydrolysis. The final equil. is not appreciably influenced by the purity of the enzyme prepn., by variations in pH , nor by the concn. of enzyme. No evidence could be obtained of the formation of an inhibitory substance (Northrop, *C. A.* 16, 1598), nor of an equil. due to the inhibitory effect of decomposition products. Addn. of another protease, e. g., that of yeast, readily carries the proteolysis to a further stage after the tryptic digestion has come to a standstill. A. W. DOX

Proteolytic enzymes in the leucocytes of man. ERIK HUSFELDT. *Z. physiol. Chem.* 194, 137-65(1931)—Different types of leucocytes were examd. separately for their proteolytic enzymes. This is the first time that such expts. have been performed with a single variety of cell, since the organs previously studied, e. g., spleen, liver and even normal blood, contain mixts. of cells. The enzyme material consisted of (1) exudates from acute inflammations where polymorphonuclear neutrophilic leucocytes are present almost exclusively (2) cells from the blood of a patient with myeloid leucemia where granulocytes and their precursors are exclusively present (3) cells from normal blood, (4) acute inflammatory exudates of rabbits which consist essentially of polymorphonuclear pseudoeosinophilic leucocytes corresponding biologically to the neutrophilic leucocytes of man. The substrates employed were casein, edestin and synthetic polypeptides, all in solns. in which the pH was varied by the aid of buffers. Proteinase action was measured by pptg. with CCl_3CO_2H and Kjeldahl detn. of the uncoagulable

N before and after the expt. Peptidase action was detd. by Willstätter's titration with 0.05 N KOH in 1% OII. In the glycerol ext. of leucocytes from the blood of a patient with myeloid leucemia, where practically all the cells belong to the myeloid system, 2 proteinases were found: a *cathepsin*, active between pH 3 and pH 7, showing optimal cleavage of casein at pH 4.3 and of edestin at pH 5.2, and a *trypsin* increasingly active from pH 4.0 on. The cathepsin resembles that of organs and represents the actual autolytic enzyme found in all animal cells. In the autolysis of cells with $CaCl_2$ the trypsin is formed in increasing quantity with progressive destruction of the cells. In the cells from acute exudates (polymorphonuclear leucocytes) trypsin is so abundantly present as to obscure the cathepsin. Attempts to adsorb the trypsin on casein did not give a satisfactory sepn., since so much of the enzyme still remained in soln. that catheptic activity could not be demonstrated. Glycerol exts. of the leucocyte mixt. from normal blood contained a considerable quantity of *peptidases*. These showed a characteristic dipeptidic cleavage like the enteric secretion. Alanilglycine was readily hydrolyzed at an optimum pH of 7.2, while the optimum cleavage of leucylglycine at pH 8.1 was only $\frac{1}{2}$ as great. Tripeptides gave the strongest cleavage, the optimum for alanylglutylglycine being at 7.3. Cleavage of the tetrapeptide, leucylglycylglycylglycine was less pronounced though distinct. In the glycerol ext. of cells from myeloid leucemia (granulocytes) the same enzymes were present and in the same relative proportions. Certain of the enzymes of the polymorphonuclear leucocytes are bound to the so-called granula of the cells. These granula are droplets of a protein and their isoelec. points can be detd. with the aid of microscopic staining reactions. In different animals the granulations have different isoelec. points. These points, as also the dissoen. of the protein, coincide with the pH activity curve of the proteinase. The isoelec. point of the pseudoeosinophilic granulations of rabbit leucocytes, pI 2.1, is at pH 7.0 and the proteinase acts from 7.0 to 3.0. The neutrophilic granulations in the polymorphonuclear leucocytes of man have an isoelec. point at 4.5, the corresponding proteinase is increasingly active from pH 4.0 into the alk. region. It is possible that the dissoen. of the protein present in the granulations detcs. the pH activity curve of the proteolytic enzyme likewise bound to the protein of the granulations.

A. W. Dox

The cephalin from human brain. II. The question of oxidation and purification of cephalin. I. I. PAGE AND MARGARET BLOW. *Z. physiol. Chem.* 194, 160-90 (1931), cf. *C. A.* 25, 718.—In contact with the air cephalin undergoes an oxidative alteration, the progress of which can be followed by noting the decrease in I no. At the same time the values for C, H and N decrease. The P content remains const. and never increases. The neutralization no. rises. A change in all these values does not occur if the cephalin is kept under 1% OII. Various preps. of cephalin show the same relationship between I no. and elementary analysis. Cephalin, like lecithin, may be oxidized by atm. O and Fe as catalyst, and the addn. of AcOH is not necessary. The product thus obtained shows the same regularity in analytical values as that obtained by atm. oxidation. Purification of cephalin with HCl by the Frankel and Neubauer procedure (*C. A.* 5, 715) leads to an ash free material, but in only 50% yield. The main losses are due to an 1% OII sol. product which does not differ from the 1% OII-insol. cephalin with respect to I and neutralization no. but gives quite different analytical values. By repptn. of purified cephalin a certain sepn. of the complex can be effected. When Et_2O - $EtOH$ is used, the component with lower I and higher neutralization no. may be centrifuged out, leaving the other component in soln. Rpptns. in the presence of H_2O lead invariably to products with lower neutralization and higher I no.

A. W. Dox

Phosphatides. II. The highly unsaturated fatty acids of the phosphatides from various organs. E. KLENK AND O. v. SCHOENBERG. *Z. physiol. Chem.* 194, 191-2 (1931), cf. *C. A.* 25, 120.—The highly unsatd. C_{27} acid recently found in the cephalin fraction of brain occurs also in the lecithin fraction of brain and in the phosphatides of beef liver. The sepn. of the highly unsatd. acids from the total fatty acids was accomplished by the Li salt-MeAc method of Tsujimoto (*C. A.* 15, 1227). The fatty acids were converted into the Me esters and fractionated in a high vacuum. The highest boiling fractions were in both cases the most unsatd. Hydrogenation and sapon. yielded behenic acid. Bromination of the unsatd. acids from liver phosphatides gave an Et_2O -insol. product, the analysis of which was in fair agreement with $C_{27}H_{44}Br_2O_2$. Removal of Br and subsequent hydrogenation, however, converted it into a mixt. contg. behenic acid, but no arachidonic acid could be identified. The highly unsatd. C_{27} acid appears to be of wide distribution in phosphatides of animal origin.

A. W. Dox

Chemistry of enzymes, hormones and vitamins. FELIX HAUROWITZ. *Pharm. Presse*, 1133-*prakt. Heft* 1930, 161-5—A review of the chief members of these 3 groups. W O E

Amylase protecting substances. IV. Protective action of infusions of yeast, malt and barley. HIROICHI NAKAMURA. *J. Soc. Chem. Ind., Japan* 34, Suppl. binding 16-7(1931), cf. *C. A.* 25, 1270—Yeast juice, prepd by plasmolyzing yeast with PhMe, then autolyzing with H_2O 24 hrs and filtering, was found to have very little protective action against the inactivation of taladiastase by heat at 50° . Infusions of malt and barley, prepd cold, then boiled and filtered, had a marked protective action but no activating action. V. Isolation of the protective substances from proteases. *Ibid.* 18-9—Sols. of papain, papayotin and pepsin, which had been found to have a protective action, were made alk. to pH 11.5 with NaOH and KH_2PO_4 and allowed to stand. The clear sols. were then found to have no protective action, but the pptd. material was 10-30 times as active as the original protease. VI. Isolation of the protective substances from peptones and malt and yeast infusions. *Ibid.* 19-20—Yeast, malt and peptone sols., treated as in part V, gave similar results, the pptd. material having a very powerful protective action. VII. The identification of the chemically pure protective substances. *Ibid.* 21-2—The ash from the ppts. obtained in V and VI was also found to have a most powerful protective action. Expts. showed the action is due entirely to the Ca content. Conclusion Amylase is an org. Ca compd and its decrease in amylolytic power when heated is due to loss of combined Ca. L E G

A unit of expression for the hemolytic index. A WEIL and D M LICHSTEIN. *Arch. ges. Physiol. (Pflüger)* 226, 239-42(1930)—Sapotoxin is suggested as a standard for expressing hemolytic activity. The hemolytic index is expressed as the no. of mg. of sapotoxin whose activity is equiv. to 100 cc. of the soln. in question. A G

The reversibility of coupled reactions in biological systems and the second law of thermodynamics. DEAN BURR. *J. Phys. Chem.* 35, 432-55(1931)—The object of this paper is to indicate the extent to which the applicability of the second law of thermodynamics to life processes has so far received direct exptl. support. In general ext. data do not exist. Two types of free-energy efficiencies are distinguished viz., "machine free-energy efficiency," in which the work done in a given reaction is compared to the total free energy dissipated by the organisms in all of their life processes, and "second law free-energy efficiency," in which the work done in a reaction is compared to the free energy dissipated by that reaction alone. Obviously, the 2 may differ greatly and it is the object to det. if the latter may ever exceed 100%. The reversibility of autotrophic reduction of CO_2 by H_2 in the case of *B. pycnoticus* (Ruhland *Jahrbuch wiss. Bot.* 63, 321 (1924)) is practically perfect, when the extraneous energy consumed in the metabolic processes of the organisms is taken into account. On the other hand the second-law efficiencies calcd. from available data on autotrophic CH_4 oxidation and heterotrophic nitrate reduction are low. Second law efficiency calcs. may be useful in detecting impossible, improbable or incomplete exptl. biological data. H F J

The specificity of animal proteases. XI. Mechanism of enzymic dipeptide cleavage. ARNOLD K. BALLS and FRANZ KÖHLER. *Ber.* 64B, 34-45(1931), cf. *C. A.* 20, 3702. 21, 2910. 23, 4235. 24, 3522—Erepsin (I) from the intestines can be sepd. into 2 components, a polypeptidase (II) and a dipeptidase (III), proving the presence in the creptic enzyme-complex of an independent dipeptide splitting enzyme. I may be used to det. the action of III in cases where the substrate is not attacked by II. The purpose of the expts. of B and K. was to det. the reaction mechanism of the enzymic dipeptide cleavage by the study of a no. of "amino peptides," many of which undergo cleavage by III. Because their constitution and properties may be varied by introducing different substituents, they are valuable for the study of the mechanism of peptide hydrolysis. The enzyme apparently attacks the peptide first at the NH_2 group, and then at the NH group. The prepn. of I, II, III and the substrates used is described by B and K. in detail. The substrates used were the following glycol derivs.: aniline (IV) *o*-toluidine, *p*-toluidine (V), *o*-aminobenzoic acid (VI) *m*-aminobenzoic acid, *p*-aminobenzoic acid (VII) *p*-nitroaniline (VIII) *p*-chloroaniline and *p*-nitroaniline-*o*-carboxylic acid (IX). Since VII, VIII and IX undergo cleavage by I, but IV, V and VI do not, B and K. conclude that the substituent group ($COOH$ or NH_2) affects the acid character of the peptide linkage and thus of the NH group. Interference with the action of III by such substances as chloroacetyltyrosine and aceturic acid supports the theory that the NH group is the 2nd point of attachment for the enzyme. The cause of the interference is the attachment of the enzyme to the NH group. The cleavage of ordinary di- and polypeptides by I, II and III was also studied. Cleavage of all of the higher glycine peptides was difficult. Apparently the presence of several

glycine residues forms an unfavorable configuration for the action of II. XXI. The causes of steric selection by enzymes. ERNST WALDSCHMIDT-LEITZ AND ARNOLD K. BALLS *Ibid* 45-8—Of the 3 isomeric chloroacetylaminobenzoic acids only the *m*-compd undergoes cleavage by the carboxy polypeptidase from hog pancreas, indicating that only in the *m* acid does the spatial arrangement of the substituents permit the reaction of the enzyme with COOH. LOUISE KELLEY

Experiments on denatured proteins. HENRY F. HOLDEN AND MAVIS FREEMAN *Australian J. Exptl Biol Med Sci* 7, 13-26(1930)—Denaturation of proteins causes an increase in $[\alpha]$ which varies with the degree of change, the treatment producing the change and its subsequent history. Ox globin, ox euglobulin, horse globin, sheep euglobulin and hen egg albumin were studied, giving the following values, resp., for native and denatured proteins, $[\alpha]_{D_{20}}$ —13° (water soln) and —72° (NaOH soln), —44.8° (NaCl) and —67.7° (NaOH), —18° (water) and —71° (NaOH), —46.6° (NaCl) and —88.8° (NaOH), —31.1 (water) and —51.4° (NaOH). The pptn by Hg salts is slight in soln not near the isoelec point, on account of a low degree of ionization. The denatured proteins including casein show a higher $[\alpha]$ in acid soln than in alkali, and when denatured by alk. a lower $[\alpha]$ than when denatured by HCl. C. G. KING

Action of dilute formaldehyde solution on proteins and protein derivatives. MAVIS FREEMAN *Australian J. Exptl Biol Med Sci* 7, 117-24(1930)—Proteins and their hydrolytic products were partially (25-70%) pptd by treatment with 0.5% HCHO at 37° for 10 days. Slight changes in $[\alpha]$ and in soly were produced, the products being readily sol in dil acids and alkalis. C. G. KING

The toxicity of blood which has been frozen. WILLIAM ALEX O'BORNE *Australian J. Exptl Biol Med Sci* 7, 223-8(1930)—Defibrinated dog blood becomes toxic to the donor when frozen and thawed and injected intravenously. Toxicity is thought to be due to a denatured protein. C. G. KING

Rabbit-liver glycogen and its preparation. MELVILLE SANJIVAN AND CARL L. ALSBERG *J Biol Chem* 89, 33-9(1930), cf C A 23, 4273—Glycogen can be readily obtained from liver by the use of 3% CCl_4CO_2H to remove the proteins. Thus prepd it contained between 0.20 and 0.18% of ash, P_2O_5 and traces of Ca and Fe were present in the ash. By hydrolysis the purity of this glycogen was shown to be from 99.4 to 99.6%; and the P content calcd as P_2O_5 was 0.032%. Glycogen is readily "denatured" by heating it to 103°; its soln is considerably less opalescent, it acquires very strong adhesive properties even to smooth glass surfaces and the port-wine color with I changes to pink as it vanishes instead of gradually fading out. A. P. LORIMER

Neutral-red reaction. V. KORPIKINO *J. Morphol Physiol* 49, 45-138(1930)—In living organisms neutral red combines with enzymes wherever they exist in digestive lumina. When cellular vacuolar systems are permeable, enzymes can usually be demonstrated upon the mitochondria. Neutral red serves as an indicator of hydrolysis and synthesis within the cell, it indicates that eggs developing in the ovary show hydrolysis in the early stages and synthesis later. Developing spermatozoa are poorly nourished. C. M. MCCAY

Enzymes. JORGE W. HOWARD *Semana med* (Buenos Aires) 1930, 11, 1028-42—The physiol and pathol action of enzymes is discussed and the therapeutic value of organ exts is emphasized. A. E. MEYER

Influence of alcohols on the heat coagulation of protein solutions buffered with acetate. TORSTEN TEORELL *Biochem Z* 229, 1-15(1930)—Heat coagulation of protein occurs only within the pH range of 4.7 to 6.2 in the presence or absence of alc. Alcs. exert an increasing inhibitory effect on the heat coagulation in the order, MeOH < EtOH < PrOH in the presence of an acetate buffer. The inhibition increases with the concn of the buffer. At an alc concn of less than 15-20% the coagulum becomes partially or completely dissolved on heating, to be repptd on cooling. Pure serum albumin or native plasma from the horse or man is similarly affected by the alc-acetate mixt., but the coagulation of the ovalbumin is not thus affected. It is pointed out that the use of alc. in tests for the Bence-Jones protein may cause erroneous conclusions. S. MORGULIS

The multiple nature of the enzyme carrier in polypeptide hydrolysis by yeast maceration juice. Preparation of protein-free eluates which act exclusively on polypeptides. A. FODOR AND L. FRANKENTHAL *Biochem Z* 229, 16-27(1930)—The kaolin adsorbate of yeast maceration juice can be leached out with an absolutely neutral soln of glycine which yields an enzyme prepn. practically free from protein; this prepn. exerts a hydrolytic action on peptone, polypeptides and dipeptides. S. M.

Hydrolysis of proline polypeptides by yeast and pancreas extracts, also by pancreatin preparations. A. FODOR, MAX FRANKEL AND S. KUK *Biochem. Z.* 229, 28-

40(1930)—The water insol residue of pancreatin preps which has a specific hydrolytic effect on dipeptides also hydrolyzes prolylalanine. The water-sol fraction which hydrolyzes dipeptides only after prolonged autolysis has no effect at first on the prolylalanine, but after 140 hrs. hydrolyzes 29%. Fresh pancreas maceration juice hydrolyzes prolylalanine, this action being increased by dialysis. The yeast maceration juice produces this hydrolysis very effectively. S MORGULIS

The effect of proteases on urease preparations. JOSEPH ZAKOWSKI *Biochem Z* 229, 41-84(1930)—The destruction of the Jack bean urease by papain and trypsin at 30° and 40° and a definite pH has been studied. At 40° with a ratio of proteases to urease 6:1 as much as 70-80% of the enzyme was destroyed in 7 hrs. Even with a papain urease mixt. in the proportion of 1:6 a loss of 52% of activity was observed under the same conditions after 7 hrs. However, it is not concluded from these observations that this indicates simply a protein hydrolysis either of the urease itself or of its carrier. The expts exclude the possibility that the inactivation of the urease is due to a liberation by the proteases of inhibiting substances. The fact that factors which activate papain and trypsin (HCN and enterokinase) also increase the destructive effect on the urease argues in favor of the enzymic nature of the inactivating process. Erepsin had no influence on the inactivation of the urease prepn. The inactivation by the various buffers employed in these expts was practically a linear function of time during 7 hrs. S MORGULIS

Effect of irradiation by means of the quartz lamps in vitro on the protein fractions of the blood plasma. LADISLAV KOSVÁK. *Biochem Z* 229, 100-8(1930)—The different protein colloids of blood manifest different sensitivities toward irradiation by quartz lamps under different disease conditions, the effect being much greater on plasma with stable than on plasma with labile proteins. The influence of this irradiation is seen on the blood plasma from normal individuals. The accelerating effect on the rate of red cell sedimentation is produced only over a limited range, beyond which the rate tends to be retarded. S MORGULIS

Some comments on the paper by Genesius on the metabolic effect of the Gurwichev's. A GLAVITCH *Biochem Z* 229, 109-14(1930)—Some important bearings of Genesius' work (*C A* 24, 5795) are emphasized. S MORGULIS

Studies on glycogenolysis. H K BARRENSCHIEEN, JOHANNES PANY and ROBERT BEACES *Biochem Z* 229, 190-215(1930)—Postmortem glycogenolysis of the liver for the first 2-4 hrs. does not proceed in a linear but in a step like fashion, the splitting off of H_2PO_4 takes place in the same way, the curves of both processes coinciding so as to suggest that the 2 reactions are coupled. As source of the new inorg phosphate there is not only the pyrophosphate fraction which is easily hydrolyzable by acid, but also the difficultly hydrolyzable fraction of the acid sol org phosphate. Neither the addn of insulin in vitro nor preliminary insulin intoxication affects the glycogenolysis or the hydrolysis of H_2PO_4 . A hexosemonophosphate ester is formed during glycogenolysis evidently at the periods when the sugar formation and hydrolysis of phosphate show a diminution. The Ba salt of the hexosemonophosphate obtained during the glycogenolysis of rabbit or dog liver is the same and corresponds to the biological or artificial hexosemonophosphate esters in reducing power and optical activity. S MORGULIS

A protein-etheral sulfate from the gastric mucosa. HARMANN MATTHIAS *Biochem Z* 229, 263-4(1930) cf *C A* 24, 4548.—Protein-etheral sulfate compds have recently been isolated from the anterior hypophysis, liver and gastric mucosa. The protein moiety of the hypophysis compd is a globulin and of its 5 S atoms 7 are in the form of H_2SO_4 and none is present as cystine. The compd prepd from liver is an acid and has no globulin properties, and of its 3 S atoms 2 are in the form of H_2SO_4 , but the third S atom is not in the form of cystine. Insulin, on the contrary, contains no H_2SO_4 but much cystine. The gastric mucosa prepn has the following characteristics. It contains only 1.67-1.77% S (liver prepn 7.49-7.68% S, anterior hypophysis prepn 2.16-2.49% S) of which 1.56% is hydrolyzable by HCl, so that the S is entirely in the form of the protein-etheral sulfate combination. S MORGULIS

Enzymes and light. XVII. The influence of the visible and ultra-violet light on the succinodehydrogenase of horse muscle. LUDWIG PINCUSSEN and WADIM ROMAN *Biochem Z* 229, 281-90(1930), cf *C A* 24, 1130.—The effect of succinodehydrogenase in the reduction oxidation system* succinic acid-fumaric acid is greatest at pH 6.9, with a second max at pH 7.7. The enzyme is more stable in an acid medium than in an alk medium. Illuminating succinodehydrogenase for 15 min with ultra-violet light causes a strong injury to the enzyme within the range of pH 6.0 to 8.0, but not so much on the acid side. The enzyme is also injured by illuminating for 1 hr with ordinary light, and it is claimed that the injury is greatest at the pH of max ac-

tivity of the enzyme. Apparently exposure to the light only for 30 min stimulates the enzyme's activity. S MORGULIS

Equilibrium between egg albumin, calcium and potassium salts. M GIUFFRÉ *Biochem Z* 229, 296-310 (1930).—Egg albumin freed by electroultrafiltration of electrolytes binds increasing amts of CaCl_2 , CaSO_4 , $\text{Ca}(\text{CSN})_2$, and of KCl from solns of increasing concn. No satn is attained up to salt concns of 0.05 *N* or 500 milliequivalents per 100 g albumin. At higher concns the behavior is irregular. Although the curves obtained for 2 different samples of similarly purified egg albumin are of the same type, the abs values for the 2 are different. Albumin binds Ca from CaCl_2 solns. of weak or medium concn more than K from KCl solns, but although in weak concn the difference may be 100%, the 2 values tend to come close together as the concns increase. In the presence of both CaCl_2 and KCl the 2 are bound. Ca from CaSO_4 is bound 10% more and from $\text{Ca}(\text{CSN})_2$ 25% less than from CaCl_2 . S MORGULIS

The postmortem lactic acid formation in the liver. IGON FENZ AND HANS LOPFER *Biochem Z* 229, 397-426 (1930).—The lactic acid formed and diffused from liver pulp upon standing 2 hrs was detd by the colorimetric method of Dische-Laszlo, with hydroquinone and H_2SO_4 as reagents. The lactic acid of the guinea pig liver increases in the first 7 hrs up to 410 mg %, but later the increase is less marked. In the human liver macerated in Ringer NaHCO_3 soln and kept in an incubator preserved with CHCl_3 , there was no further increase in lactic acid after 7 hrs and therefore the livers obtained 8 hrs after death already attained the max lactic acid production. In some 60 cadavers the av amt of lactic acid was about 0.30% (0.136 to 0.574%), and since among these were persons with diabetes, cirrhosis, acute yellow atrophy and various endocrine disturbances there was thus no evidence of any dependence upon the total glucide content of the liver. Also in livers which 10-33 hrs after death remained for many hrs without CHCl_3 at 37° either in NaHCO_3 or in Na_2HPO_4 , even with the addn of glucose, there was no further increase in the lactic acid. Expts on the livers from guinea pigs show that shortly after death the lactic acid formation is considerable, becoming somewhat less after 2 hrs and very much less after 4 hrs, almost ceasing after 7 hrs. There was no difference whether NaHCO_3 or Na_2HPO_4 was used in the Ringer soln. Nor could the addn of glucose, levulose or of glycogen cause any alteration, and only occasionally was there an irregular increase in the lactic acid upon the addn of hexosediphosphate. Otherwise the max value of 500 mg % was never exceeded. In animals with a particularly low glycogen content of the liver the max formation of lactic acid fell below this value. Similar results were obtained in studies on the kidneys. The results of the lactic acid formation in practically all organs of 2 human cadavers and of 2 guinea pigs are tabulated. S MORGULIS

The cleavage of tyrosine and tryptophan from casein by papain activated by hydrocyanic acid. THEODOR LEIPERT AND IRMGARD HAFNER *Biochem Z* 229, 427-32 (1930).—Only 47% of the total *N* of casein is hydrolyzed as amino *N* in 12 days' digestion by papain-HCN. Tyrosine is set free entirely in proportion to the total amino acid cleavage, and tryptophan is only more or less proportional to the total hydrolysis. Under the combined action of papain-HCN and trypsin for a week 71% of the total *N* is hydrolyzed as amino *N*, but the hydrolysis of tyrosine and tryptophan by the 2 enzymes together seems to be unaffected. S MORGULIS

Studies on tobacco enzymes. CARL NEUBERG AND MARIA KOBEL *Biochem Z* 229, 455-63 (1930).—In the present day methods of prepn roof-dried tobacco is passed through a slume with a gradually rising temp, for the purposes of removing the H_2O . In the zone of max heat the temp reaches 100°. The entire passage occupies 60 min. It was found that the tobacco enzymes do not lose their activity in this process. The following enzymes were studied: amylase, invertase, phosphatase, pectase, glycolase and ketonaldehydemutase. S MORGULIS

A model for studying pectase. CARL NEUBERG AND CLARA OSTENDOERF *Biochem Z* 229, 464-6 (1930).—The action of pectase may be represented as follows: Under the influence of this hydrolyzing enzyme upon probably the acid Me ester of a complex polybasic glucidic acid (pectin), the MeOH is split off while the poorly sol. Ca and Mg salts of the pectic acid form a gelatinous ppt. For better investigation of the action of the pectase a substrate of definite chem composition, namely, the Ca salt of methyl-d tartaric acid, is substituted for the pectin whose compn is highly variable. This substance under the influence of pectase gives MeOH and the sparingly sol Ca tartrate. S MORGULIS

Studies in comparative biochemistry. VII. Chemistry of the special saké yeast. JUN-ICHIRO SAGARA *J. Biochem. (Japan)* 12, 459-71 (1930), cf *C. A.* 20, 3745.—In 2 strains of saké yeast there was a close chem. resemblance both in the inorg. and

org compn No lecithin was found but a phosphatide was bound, which because of its high N content must be either a tri- or diamino-monophosphatide. The presence of a pentose was proved, and a substance related to bioses was isolated in pure cryst form. Arginine, histidine, lysine, phenylalanine, proline and valine were identified, the phenyl alanine and proline being the predominant components of the protein. The nucleic acids of both strains are distinguished by their low N and P contents. Both types of yeast contain amylase, esterase, glycerophosphatase, pepsin and trypsin. Data are also recorded on the inorg compn. S. MORGULIS

The properties of the biochemical constituents, the proteins in particular, in anhydrous solution. JEAN LOISELEUR *Compt rend* 191, 1477-9(1930)—In aliphatic solvents, biochemical substances form true, not colloidal, solns. They manifest certain chem. properties: (1) they are not frothy, their viscosity is less than aq solns of the same concn, they are optically clear, not opid by a formic soln of the ions Na, K, Mg, Fe, Cr or by tannin, MeOH, CHCl_3 , pyridine or cyclohexanol (2) Addn of water to an aliphatic soln of protein forms a colloidal dispersion which shows a Tyndall effect, the micellar state, adsorption, and flocculation (3) Casein and keratin, in acetoformic soln and in presence of HCl and HNO_3 , form a nitration deriv having colloidal properties. N. M. NAYLOR

Growth factor in plants. HANS A. FULER, MARGARETA RYDBOM AND HARRY HELLSTRÖM *Svensk Kem Tids* 42, 277-81(1930)(in German)—Petr ether exts. of sprouts examd spectrographically showed bands at 533, 497, 447 and 424. Distd at 150° and 0.02 mm Hg they gave bands at 256 and 325. A distillate from a carotenoid from swordfish liver had a band at 325 and Moore reports 328 for a rat liver fraction (C. A. 24, 5343). A. R. ROSS

Membrane phenomena in living matter, equilibrium or steady state. A. V. HILL. *Trans. Faraday Soc* 26, 667-78(1930)—The difference in l. p. between the white and yolk of hen's eggs, observed by Straub (C. A. 23, 3482), is confirmed. Since it persists in an atm of H the oxidative mechanism at the membrane postulated by Straub for opposing the osmotic pressure cannot occur. The p. d. between inside and outside of muscle or nerve is also far greater than can be accounted for on the Donnan equilibrium basis. The widespread occurrence of such steady states, maintained by the expenditure of energy, is held to indicate their biol. importance. K. V. THIMANN

The state of water in colloidal and living systems. ROSS AIKEN GORTNER *Trans. Faraday Soc* 26, 678-704(1930)—The work of Newton and Gortner (C. A. 18, 2543 3205) and that of Robinson (C. A. 23, 910, 1180) on winter hardiness are reviewed in relation to the general problem of "bound" water in the cell, and are compared to similar phenomena with inorg gels and proteins. A discussion is also given. K. V. THIMANN

The *pu*-stability regions of the proteins. T. SVEDBERG *Trans. Faraday Soc* 26, 740-4(1930), cf C. A. 25, 530—Electrophoresis measurements show that the isoelec pt. of proteins is not identical with the point of min. acid and alkali binding, but is identical with the soly min. The method can be used to sep mixts of proteins, since mutual neutralization does not occur. The isoelec pts. detd in 0.02 N NaOAc buffer are: ovalbumin 4.55, serum albumin 4.88, R phycoerythrin 4.25, C phycoerythrin 4.76, hemocyanin (*Helix*) 5.05, Bence Jones protein 5.20. A summary of the results of sedimentation detns is given. K. V. THIMANN

The solubility of proteins. E. W. J. MARDLEY *Trans. Faraday Soc* 26, 750-60(1930)—The soly in water of Coignet's gelatin varies with the amt. of the solid phase, and is greater when detd by cooling soln than by soaking gelatin in cold water. It increases sharply at 26° , is diminished by washing, and is increased by the presence of HOAc, HCONH_2 , urea and other org compds. It is concluded that gelatin is heterogeneous (cf Schryver and Thimann, C. A. 22, 963), from the corresponding behavior of blood albumin a similar conclusion is drawn. K. V. THIMANN

Biological significance of the physical state of lyophilic colloids. F. P. NORD *Trans. Faraday Soc* 26, 760-8(1930)—A summary of the work (cf C. A. 22, 4544, 23, 5475) in which increase of zymase activity on treatment with C_2H_2 and C_2H_4 is correlated with increase of surface tension and decrease in viscosity of cell free zymase solns. Corresponding surface tension and viscosity phenomena with solns of albumin or Na oleate indicate surface adsorption of these gases, and also of N_2O , on colloids, leading to protective action and hence increase of enzymic activity. K. V. T.

Phase rule equilibria of horse serum globulin. JAMES W. McBAIN AND ELOISE JAMISON *Trans. Faraday Soc* 26, 768-9(1930)—The phase diagram for globulin prep'd by Svedberg's method water and Am_2SO_4 at *pu* 6.0 at 40° is given. "It appears that globulin solns, euglobulin and pseudoglobulin are but 3 phase of a system

of the same mother substance, dehydrated globulin—sols of globulin being the ordinary isotropic soln, while englobulin and pseudoglobulin are liquid crystal phases, or a liquid and glass which are slightly doubly refracting." K. V. THIMANN

The structure, behavior and physiological characteristics of vertebrate cells cultivated in vitro. H. B. FELL AND E. N. WILLMER *Trans Faraday Soc* 26, 772-9 (1930) —A general lecture on the technique of tissue cultures and their physical behavior. K. V. THIMANN

Surface structure in the integration of cell activity. RUDOLPH A. PRINGS *Trans Faraday Soc* 26, 797-807 (1930) —The importance of interface reactions in determining the activities of the cell is discussed. The interfacial tensions of palmitic, capric and erucic acids in benzene solns against phosphate buffer show a sharp fall at about pH 6 continuing to pH 9.5. Adsorption of weak acids and bases on purified charcoal follows closely the dissociation curves, the undissociated mol being the most adsorbed (Phipps and Peters, *C. A.* 23, 5080). The pptn of phosphotungstates of histamine and other bases is also a function of pH , histamine being pptd sharply at pH 7.8 and on the acid side. The presence or absence of a single H^+ ion is called to be of considerable importance in determining the reactions occurring in bacterial cells. A theory of mosaic cell structure, involving mobile chains of liquid crystals a few mols thick, and orientating all substances in the cell, is developed. K. V. THIMANN

The isoelectric point of cells and tissues. HANS L'PPIFFER *Trans Faraday Soc* 26, 822-36 (1930), cf *Protoplasma* 11, 85 (1930) *Biol Reviews* 4, 1-40 (1929), and *C. A.* 23, 3725. K. V. THIMANN

Decrease in the activity of β -glucosidase of emulsin of almonds in the course of successive syntheses of β -methylglucoside. M. BRIDEL AND N. JOANNIS *J. pharm. chim.* [8], 12, 337-9 (1930) —A mixt of 3 g emulsin and 150 g dextrose in 500 cc of 70% MeOH was continuously rotated for 8 days, *s. e.*, until the optical rotation of the soln. became const ($42^{\circ}40'$). The emulsin was then collected, washed with 90% alc and dried *in vacuo*. Only 2 g was obtained. The process was repeated 6 times with the recovered quantity of emulsin and 60 times its wt. of dextrose in each case. The following wts of emulsin were recovered: 1.2, 1.2, 1.0, 1.0, 1.0, 1.0 g. The 7th synthesis required 19 days to reach equil., and in each of the 7 cases, about 77% of the dextrose became combined. During equil periods of time in the 7 expts, the activity of β -glucosidase steadily diminished. Thus, measured by time, emulsin lost more than half of its activity. Nevertheless, having acted in these expts for 3 months on 350 times its wt. of dextrose to cause 77% of it to form β -methylglucoside, it was still active enough to complete the last synthesis in 19 days. Mech rotation shortened the time of these expts materially. Also in *Bull. soc. chim. biol.* 12, 931-3 (1930). S. WALDBOTT

Hydrolysis by emulsin of 2 glucosides not considered hydrolyzable by this enzyme: asebotoside (asebotin) and phlorhizoside (phlorhizin). M. BRIDEL *J. pharm. chim.* [8], 12, 385-93 (1930) —Contrary to the accepted belief, asebotoside (cf Bourquelot and Tschentenholz, *C. A.* 6, 1207) and phlorhizoside (cf Johnson and Robertson, *C. A.* 24, 1805), which are β -rotatory glucosides, are hydrolyzable by emulsin of almonds, yielding in each case d glucose and a nonglucidic product insol in water. Both glucosides therefore belong to the class of β -glucosides, *s. e.*, they are hydrolyzable by β -glucosidase. Twenty five references are given. Also in *Bull. soc. chim. biol.* 12, 921-30 (1930). S. WALDBOTT

The denaturation of albumin (BANCROFT, RUTZLER) 2. The influence of the wave length of the exciting rays on the fluorescent spectrum of etioporphyrin (AHARONI, DUBOS) 3. Denaturing of proteins of soy beans by alcoholic extraction (OKANO, *et al.*) 27. The influence of light, CO and quinone on the methylene blue reaction (TAMURA, *et al.*) 10. The bile acids (WIELAND, *et al.*) 10.

Keijo Journal of Medicine (*New journal*). Published about quarterly by the Medical Faculty of the Keijo Imperial University, Keijo, Chosen, Japan. Vol 1, No 1, appeared Feb 28, 1930. This journal replaces *Acta Medica et Chirurgica in Keijo* which was discontinued at the end of Vol 12, 1929.

B—METHODS AND APPARATUS

STANLEY R. BENEDICT

The antineuritic vitamin (preliminary communication). A. G. VAN VEEN, *Rec. trav. chim.* 49, 1178-9 (1930) —According to the method of Jansen and Donath (*Proc. Acad. Sci. Amsterdam* 29, 1300 (1920); cf *C. A.* 20, 2005) a product which contains about

20% of the vitamin may be easily obtained, the further purification being, however, very difficult and giving only 50-100 mg of the pure vitamin from 100 kg rice polishings after a few months' labor. Attempts have been made by Guha and Drummond (*C. A.* 24, 1140) and by Williams, Waterman and Gurn (*C. A.* 24, 4535) to obtain larger yields but without success. The technique of Jansen and Donath for the purification of the 20% product had to be modified profoundly in order to get higher yields. After having studied the products which accompany the vitamin in the 20% product, the author has now succeeded in getting 140 mg of the cryst vitamin from 75 kg rice polishings, this product was, however, not pure, but could be purified by recrystallization. The 20% product may be freed from several physiologically active compds by treatment with acid chlorides or, even better, with sulfonyl chlorides, which leave the vitamin itself unchanged, and this may be purified further with the double salts of Pt or Cd chloride. It was found also that the Pauly reaction with diazotized sulfanilic acid is fainter the purer the vitamin. This fact was confirmed by Jansen (private communication). C F VAN DUIN

An improved design of the Van Slyke apparatus for the estimation of amino nitrogen SIDNEY J POLLEY *Biochem J* 24, 961-4(1930) BENJAMIN HAWRO

An apparatus for rotating serum-antigen mixtures in the Kline test for syphilis PHILIP L VARNEY *Am J Pub Health* 21, 65-9(1931) —An electrically driven device is described, which mechanically mixes by rotation the serum-antigen mixt. As many RS Ø slides (2" X 3") are accommodated at one time. The rotary speed of the app has been arbitrarily set at 120 r p m through an arc of 1/4" radius. The app may be connected to an elec timing clock, thus relieving the technician of the necessity for timing the period of rotation resulting in a greater uniformity of results and a saving in time necessary to perform the tests J. A. KENNEDY

Colorimetric method for determining silicon. EARL J KING *Bull soc chim Mol* 12, 903-9(1930), cf *C. A.* 23, 405, 25, 602 —The phosphate error in Isaac's method (*C. A.* 18, 3199) is pointed out and a modification proposed to avoid it C G K

The estimation of copper in organs. EMILE CHERBULIEZ AND STEFAN ANSBACHER *Arch Path Anat (Virchow's)* 278, 365-71(1930) See *C. A.* 24, 2402 J R L

A new microcolorimeter. A G SNEPTEL *J Lab Clin Med* 15, 1016-8(1930) E R MAIN

A graduated test tube colorimeter ARTHUR T BRICE, JR *J Lab Clin Med* 15, 1018-22(1930) E R MAIN

Microdetermination of urea by Nicloux and Welter's method (GUILLET, GOLAZ) 7 An improved pipet manipulator (*Duff*) 1 A new high-efficiency glass condenser (CRANDALL) 1

Handbuch der biologischen Arbeitsmethoden. Edited by EMIL ABERMOLDEN Abt 1 Chem Methoden Tl 2 Allgemeine chem Methoden Hälfte 2 H 6 (Lfg 330) FRANZ BACHER Die Verseifung Pp 2385-2378 Reviewed in *J Am Chem Soc* 53, 818(1931) M 16 H 7 (Lfg 339) HANS HORSTERS AND HELENE HORSTERS Neuere Synthesen biologisch wichtiger Pyridinkörper, Pp 2879-2970 M 5 Abt 3 Physikal-chem Methoden Tl A Allg u spez Methoden zur Untersuchung d Verhaltens gelöster Stoffe H 10 (Lfg 342) GÜNTHER RIENÄCKER Potentiometrische Mikrotitration. FRANZ LEUTHARDT Pufferung und Puffersysteme ROBERT BRINKMAN Registrierung d Wasserstoffionenkonzentration im strömenden Blut, HANS FISCHER AND OTTO REHMANN Die Bestimmung von Gas- und Dampfdrücken Pp 1913-2201 M 17 Abt 4 Angewandte chem und physikal Methoden Tl 14 Untersuchung d Nahrungs und Genussmittel H 3 (Lfg 340) WILHELM PLÜCKER Allgemeine Methoden zur Untersuchung der Nahrungs- und Genussmittel Pp 721 1034 M 17 Cf *C. A.* 24, 5391 Berlin and Vienna Urban & Schwarzenberg

Polarimeter for testing urine OPTISCH-MECHANISCHE FABRIK STEINDORFF & CO Ger 485 094, Jan 21, 1928

C—BACTERIOLOGY

CHARLES B MORREY

Vital conditions of ferruginous bacteria. (MILLER) I TOKOWSKA *Bull intern acad polonaise* 1929B I, 255-62 —Fe bacteria are found in certain wells and other waters of Poland in which the Fe content is usually greater than 2 mg per l and the reaction is between pH 5.88 and 7.60. Large concns of salts appear to be unfavorable. The bacteria are able to withstand temps. down to 0°. B C A

Disinfecting properties of chloroxym. A. SWENSON. *Z. Desinfekt.* 21, 109-121 (1920). Chloroxym has a high penetrating power for porous materials. Its disinfecting power at 50° is double that at 4°. Spore-forming bacteria are very resistant. B. C. A.

Studies on the Coccaceae. XVI. Genus *Leuconostoc*. G. J. HICKER AND CARL S. PETERSON. N. Y. Agr. Expt. Sta., Tech. Bull. 167, 584 (1920). Of C. 23, (20) *Leuconostoc* study was made of 166 strains of low acid producing types of streptococci (*Leuconostoc*) from dairy sugar beets, fermenting vegetables and from milk and milk products. All members of the genus *Leuconostoc* produce approx. 45% lactic acid from glucose, 20% CO₂ and 25% of volatile products, including AcOH and H₂O. In addition, the members of this group produce mannitol from fructose and a levulose or dextran from sucrose. In morphology, the genus *Leuconostoc* lies intermediate between the streptococci and the lactobacilli. Secondary characters used for recognition of species within the genus are the fermentations of pentoses and sucrose. Three species are well defined: *L. mesenteroides* ferments pentoses (arabinose or xylitol) and sucrose, produces slime in sugar beets, and is generally found in fermenting vegetable material and in sugar beets. *L. zeiniformis* ferments sucrose but not pentoses, produces a moderate amt. of slime in sugar beets and may be associated with either vegetable or with dairy products. *L. citreum* fails to ferment either sucrose or the pentoses, produces no slime in sugar beets and is generally associated with milk or milk products. C. R. FILLERS.

Staining bacterial flagella. R. P. LEONARD. Sci. Rep. 11, 272-7 (1920). Arnold fed Zetsohn technique is preferred. C. R. FILLERS.

The physiology of the acetic acid fermentation. I. The importance of cytochrome in the physiology of cell respiration. HIROMI TANIYA AND KIYOSHI TANAKA. *J. Sci. Planting*, 5, 117-111 (1920).—The AcOH fermentation with *Acetobacter pasteurianus* was studied in a special app. It consists of 2 cylinders carefully ground into each other providing a piston-like airtight closure. The outside cylinder is provided with an opening and a two-way stopcock. A series of 10 such cylinders is gently shaken on a machine. By this arrangement samples of both the gas and the fermentation must can be taken. It was found that quinone can also be used as a H₂ acceptor for the AcOH fermentation instead of O₂. The extent of the fermentation is largely independent of the amt. of O₂. CO₂ does not influence the quinone fermentation while it hinders noticeably the O₂ fermentation. In the presence of light this effect is markedly ruled out. The indophenol reaction of the AcOH bacterium is also hindered by CO₂ more so in the dark than in the light. Toluene was found to retard strongly the O₂ fermentation even when the O₂ amts. were large; the quinone and methylene blue fermentation, however, is not influenced. While the indophenol reaction of the AcOH bacterium as well as yeast is also retarded by toluene, *Leuconostoc* seems to be indifferent to its action. It furthermore was shown that the normal functioning of cytochrome is disturbed by the addition of toluene and quinone, but is resistant toward methylene blue. The phenomena encountered are explained by assuming that cytochrome plays an important role in the O₂ fermentation of the AcOH bacterium and the normal O₂ respiration. Cytochrome is the "O₂ pressure regulator." The equation $Q = K' C_A$ (Q = extent of fermentation, K' = activity of the oxidizing enzyme system, C_A = concn. of the acceptor substance present on the surface of the enzyme) is given and the process discussed theoretically. The expt. findings seem to agree with the theory of K. Saito and H. Taniya (C. A. 24, 3337). B. C. F.

D—BOTANY

THOMAS G. PHILLIPS

Peroxiside of black willow bark. M. BRIDEL AND J. RABAT. *Fall. an. chim. Soc.* 12, 332-41 (1920).—The extn. of the glycoside (see C. A. 24, 4070) is described. B. C. A.

The action of radium on the germination of seeds. G. MESSADROLI AND E. VARRONE. *Atti. Acad. Lincei* 12, 77-84 (1920), cf. C. A. 24, 5535.—A variety of seeds (cereals such as oats and corn; legumes such as peas, beans, soy) were exposed to the action of the same quantity of Ra at the same distance for 5 min., 30 min. and 3 hrs. The seeds were then planted and the relative growth was noted. Five min. is enough to stimulate growth somewhat; 30 min. has the max. effect, while with 3 hrs., growth is retarded. If 1/2 the amt. of Ra was used, 5 times the length of exposure was necessary. Therefore, Ra has a stimulating effect up to a certain period of exposure, beyond which the seeds are harmed. The stimulation is apparently permanent as seeds planted 2 months

after exposure showed the same accelerated growth as seeds freshly exposed for the same period of time.

A. W. CONTIERI
Bibliography of *Aspergillus*, 1729 to 1928. HIROSHI TAMURA AND SHUNKICHI MORITA. *Botanical Mag. (Tokyo)* 43, No 106 to 44, No 124, pp 204(1929-30); cf C. A. 24, 2496.—A chronological bibliography of 2424 papers dealing with the taxonomy, morphology, physiology and enzymic and chemical activities of *Aspergillus* species, with author index.

A. W. DOX
Researches in ionic interchange between yeast cells and salt solutions. PAUL GENAUD. *Ann. physiol. physicochim. biol.* 6, 240-330(1930), cf C. A. 25, 316.—When yeast cells are placed in solutions of dissociable salts, there is a rapid interchange of cations within a few min. between the cellular membrane and the surrounding fluid, the metal begins to penetrate the cell interior within about 10 min., complete equilibrium being reached in about 60 min. Vital staining by basic dyes is due to just such a cationic interchange between the membrane and the surrounding fluid. The total cation content is a fairly constant quantity, characteristic for a yeast cell under definite conditions of growth. Although the cationic interchange usually concerns equimol quantities, with some metals (e.g., Ag) there is an enormous absorption not conditioned by a simple interchange, but by other forces (e.g., reduction of the Ag^+ to Ag). In cases of simple diffusion (K^+ , NH_4^+ , Pb^{++}) the vacuoles contain approx. half as much of the absorbed cation as the membrane. Dead yeast cells can take up 4 times as much Pb as living cells, probably because of the liberation of free acids. Corresponding to their slight Na^+ content, the cells are impermeable to Na^+ , possibly on account of the fact that the hydrated ion cannot pass through the pores of the cell membrane. The enormous avidity of the cell for Ag^+ is conditioned by the fact that the latter is reduced by an as yet uncharacterized substance in the cell interior. There is an extensive bibliography. H. E.

Effect of spiral ringing on solute translocation and the structure of the regenerated tissues of the apple. L. H. MACDANIELS AND OTIS F. CURTIS. Cornell Agr. Expt. Sta., *Mem.* 133, 3-31(1931).—Trunks of young apple trees were ringed by removing narrow strips of tissue extending twice around the trunks spirally. In some cases the phloem and outer tissues only were removed, but in others the outer annual ring of xylem also was cut out. The effect of this ringing upon solute translocation was studied by making N analyses and catalase detn. on the leaves. The structure of the regenerated tissues was studied by sectioning and microscopic examn. Lateral transfer of solutes and of food substances is relatively slow, though it does occur when forced by spiral ringing. Foods from above and nutrient salts from below tend to move in straight lines parallel to the long axes of the conducting elements. When the trunk is spirally ringed and nitrate is added to the soil, N movement to branches immediately above the end of the spiral is greatly reduced as in a completely ringed trunk, whereas movement to those branches on the side not obstructed by this upper end of the spiral is normal as in the check trees. The phloem is the more important tissue involved in translocation, since the effect of cutting the phloem alone was the same as when both xylem and phloem were cut. There is a close correlation between high N content and high catalase activity of the leaf tissues of all trees examd. The conducting tissues formed by the cambium subsequent to the ringing changed their orientation so that the long axes of the elements were parallel to the spiral. After the new tissues were formed, the localized effects of the ringing upon solute transfer to the leaves on the tree top disappeared. Cambial growth is made in response to the coming together of the food from the leaves and the nutrients from the roots. The position of greatest thickening of the xylem is detd. more by the food supply than by the nutrient salts. The change in the orientation of the cambium is detd. by a change in the direction of the movement of foods and of nutrients, or possibly something of the nature of an elec. polarity may be a factor in the orientation. The general conception that lateral conduction of food and of nutrients is slow, that these substances move in straight lines parallel to the long axes of the vascular elements, and that cambial growth is a response to the coming together of the streams of food and nutrients is valuable in explaining plant response in practical problems of propagation, grafting and pruning. The possible bearing of this conception of greatest growth activity upon the practice of tapping trees for oleoresin is discussed.

C. R. FELLERS
Growth behavior and maintenance of organic foods in Bahia grass. W. A. LEUKEL AND J. M. COLEMAN. Fla. Agr. Expt. Sta., *Tech. Bull.* 219, 5-56(1930).—Except for the first year, frequent cuttings yielded more top growth and greater wt. than a single cutting at mature growth. The top growth from plants cut frequently retained more of a vegetative growth condition, more uniform % of N and a narrower carbohydrate:N ratio than plants grown to maturity. The latter gradually decreased

in % N, were less vegetative and gradually showed a wide carbohydrate-N relation as associated with increased reproduction. Increased vegetative top growth production on plants was associated with a variation in % and quantity of org foods in the stolons. During periods of slow growth such stolons again increased in wt and quantity of reserve foods. Heavy seed production was associated with some decrease in org foods in stolons but such foods again increased in quantity thereafter. The elaboration of org foods by the more horizontal leaf area not removed by cutting appeared to be sufficient for the growing needs of the plant. C. R. FELLERS

Carbohydrate metabolism of *Stipa pulchra*. ARTHUR W. SIMMONS and EDWARD C. MCCARTY. *Hilgardia* 5, 61-100 (1930).—The growth cycle of *Stipa pulchra* is an orderly process characterized by periodicity in that intervals of rapid growth alternate with intervals of depression in the growth rate. An inverse correlation existed between the annual march of carbohydrates and the growth rate. Accumulation of foods is related to low or to declining growth velocity and is most active near the close of the annual growth cycle. Practically complete accumulation of carbohydrate foods occurred where 43-50% of the total annual herbage yield was produced prior to the peak in the growth rate. Clipping or grazing had little effect on the total herbage yield and did not prevent the accumulation of max amts of carbohydrates in the latter part of the growth cycle. The herbage of *Stipa pulchra* on this range retains its succulency for 9-10 months. C. R. FELLERS

Experiments in forcing gladioli. F. F. WEINARD and S. W. DICKER. Ill. Agr. Expt. Sta., *Bull.* 357, 345-61 (1930).—The most desirable forms for forcing are those over 1 1/2 in. in diam. Storage at 5-8° early in the season and at 20-1° for a short time previous to planting gave the best results. Ethylene dichloride and ethylene chlorohydrin forced the corms into early growth. The chem. treatment, however, is not yet well standardized and sometimes causes injury. C. R. FELLERS

Some effects of defoliation on fruit spur composition and fruit bud formation in the Oldenburg apple. G. I. POTTER, H. R. KRAVILL, S. W. WENTWORTH, J. T. SULLIVAN and P. T. BLOOD. N. H. Agr. Expt. Sta., *Tech. Bull.* 41, 3-27 (1930).—All the blossom buds were removed when in the "pink" stage from 3 heavily blooming Oldenburg apple trees. Every second spur on 2 similar trees was defoliated so that along each branch defoliated spurs alternated with those blossoming and setting fruit. Four trees from which no blooms were removed served as controls. The wt and the compn. of the developing fruits were detd. at 10 dates between the period of petal fall and harvest. The compn. and the fruit bud formation of bearing spurs from the control trees and of defoliated spurs from the other 2 treatments were studied. The increase in dry wt. of the individual fruits followed the usual logarithmic curve of growth, the rate of increase reaching a max. in the second and third weeks in July. At this period each apple absorbed about 1 mg. of N daily and about 300 mg. of carbohydrate. Throughout June and July the N content of the bearing spurs was much higher than that of either of the other 2 types, and also was considerably higher on a % basis than the N content of the fruits. The starch content of bearing spurs and the adjacent 50% defoliated spurs was similar and lower throughout the season than in the 100% defoliated spurs on non fruiting trees. The sugar starch equil. of the fruit was independent of that in the spur to which it was attached, and sugars in the fruit reached a concn. several times that in the spur. The ash content of the bearing spurs was comparatively low, that of the 2 types of defoliated spurs was essentially similar. No difference large enough to be detd. existed among the 3 types of spurs in content of free reducing substances, sucrose, P, phlorhizin, acid hydrolyzable substances and total carbohydrates. The compn. and fruit-bud formation of adjoining spurs are materially altered by the presence of blossoming and fruiting spurs. No critical evidence is afforded, however, as to the fundamental nutritional causes underlying fruit bud formation. C. R. FELLERS

Relation of hydrophilic colloids to hardness in the apple as shown by the dye absorption test. STUART DUNN. N. H. Agr. Expt. Sta., *Tech. Bull.* 44, 3-18 (1930).—Results are compared both to general experience and opinion as to the ranking of varieties and to a freezing test by solid CO₂. Results of tests for 5 yrs. on 1-yr.-old twigs gathered each fall for 22 different varieties grown in Iowa and of direct freezing tests for 2 of these yrs. show a correlation of about 65% of the varieties compared to field experience, and 55% in 1 series and about 80% in the other on the basis of the parallel freezing test. The freezing tests agree with each other to a less extent. A marked consistency in the varieties not in agreement, as shown by a comparison of the various tests, and their variations in position in the different series, suggest that a lack of constancy in hardness in any variety is associated with its failure to respond consistently or correctly.

to any hardness test. A similar conclusion is drawn from results on N. H. varieties. Similar tests on 12 N. H. varieties show a correlation between the dye test and either of the other standards of from 60 to over 90%. A large no. of freezing tests were made, distributed throughout the dormant season, which offered a wider basis of comparison. A marked trend toward a closer agreement was shown in the series gathered later in the season. Conclusion: The dye absorption test is not sufficiently consistent to be used in a quick, conclusive detn. of hardness in the apple, but possibly in conjunction with other tests, such as the freezing test, will serve to differentiate between the hardness of unknown varieties which are rather close in that respect, if a sufficient no. of detns. is made. Some other factors besides the hydrophilic colloids are apparently partially responsible for hardness, at least in ligneous plants like the apple. C. R. F.

Development and ripening of peaches as correlated with physical characteristics, chemical composition and histological structure of the fruit flesh. II. Histology and microchemistry. P. M. ANTON, G. T. NIGHTINGALE AND M. A. BLAKE. N. J. Agr. Expt. Sta., Bull. 507, 3-19 (1930), cf. C. A. 24, 5223—Elberta and Shipper Cling varieties were used. Microchem. tests showed a gradual decrease in *pectin*, *cellulose* and *hemicellulose* content as ripening proceeded. The change from hard ripe to soft ripe occurs within a very short time and is accompanied by profound changes in the cell walls, many of which completely break down and become disorganized. Comparisons made on fruits from a highly vegetative tree and a less vegetative tree, higher in carbohydrates, showed no chem. differences for the first 75 days after bloom. However, the latter contained more tannins and oxidases, but no histological differences were noted. Differences in microchem. and histological reactions in the Elberta and Shipper Cling varieties were evident only in the final stages of ripening. The non-ripening quality of ripe fruits of Shipper Cling variety is ascribed chiefly with the retention of a high content of insol. *pectin* in the cell walls of the flesh. A bibliography is given. C. R. FELLERS.

Studies on tree root respiration. III. G. H. HARRIS. Sci. Agr. 11, 191-9 (1930), cf. C. A. 24, 5551—The immediate influence of top pruning on root respiration is due to the injury caused. On the other hand, the immediate influence of injecting sugar in the top on root respiration is due to the injections stimulating bud activity. The buds in turn stimulate root respiration. In the nutrient waters used, where all the nutrients were readily available, the amt. of CO_2 excreted by the root and absorption appeared to have little relation. Absorption seemed more closely related to the activity of the growing top. Cutting off the light from the growing tops reduced the amt. of root respiration. When transpiration is high, the CO_2 excreted by the roots is lower than if transpiration were low at this time. Some of the CO_2 produced during root respiration is apparently swept up to the tops during a time of high transpiration and consequently less is excreted by the root than if transpiration were low. C. H. F.

Sparasole in the roots of *Rhododendron japonicum*. ROYO KASUYITA. Ada Phytochem. 5, 157-65 (1930).—Sparasole (I) is obtained from the alc. exts. of roots of *Rhododendron japonicum* collected in November. I is found in the oxid. recover upon distn. of the alc. It m. 67-68°. The yield is 0.25%. Hydrolysis yielded the monomethyl ether of oxalic acid (II), m. 174-175° from 80% alc. The m. ps. given by previous authors are too low. II is converted to oxime by osmid. HCl or HI . A mixed m. p. with a sample of I isolated from another source finally identified the substance as the monomethyl ether of oxalic acid methyl ester. F. G. F.

The amount of sorbitol in rowan berries. G. TAYLOR. Bull. soc. chim. bel. 12, 741-3 (1930).—The wt. % (based as the diacetaldehyde deriv.) found on Aug. 11, Sept. 12 and Oct. 17 was 44.4, 66.5 and 68.9 g. per kg. whole berries of *Sorbus aucuparia* corresponding to 147.5, 200.4 and 173.2 g. total sd. material. Acid and sugar analyses are also given. C. G. KRIG.

The penetration of 1-naphthol-2-sulfonic acid in daphenol, o-chlorophenol indophenol and o-cresol indophenol into Valonia. MATHILDA MÖLLENTHALER-BRINKS. Proc. Natl. Acad. Sci. 17, 1-6 (1931).—B. (C. A. 23, 4241) prevents the continuation of the studies on oxidation-reduction indicators and their penetration into living plants of Valonia. Indicators lying further toward the pos. end of the scale are considered.

Naphthol-2-sulfonic acid indophenol does not penetrate into the sap within 24 hrs. at pH values from 5.8 to 9. o-Cresol indophenol and o-chlorophenol indophenol penetrate into the sap in a colorless form and faster from an acid than alk. soln. Different equil. concns. are estab. according as a phosphate or borate buffer is used. Where reduction has changed the color of the dye, its identity was estab. by spectrophotometry. B. S. LEVINE.

A source of error in the determination of amide nitrogen in plant extracts. HERBERT

B VICKERY AND GEO W. PUCHER. *J Biol Chem* 90, 179-88(1931).—HCl is an unsuitable reagent for the hydrolysis of amides in plant extracts and may lead to uncertain results if nitrates are also present, although the so-called Sachsse method in which NH_4Cl is used and hydrolysis is continued for 4 hrs. actually gives quite accurate results if these definite conditions of hydrolysis are rigidly maintained. A compensation of errors occurs in which the oxidation of NH_4 by the mixt. of HNO_3 and HCl is counterbalanced by extra NH_4 formed by reduction of HNO_3 under the influence of an easily oxidized unknown substance present in tobacco leaf and other plant extracts. Substitution of 2 $N H_2SO_4$ and a 6-hr. hydrolysis are recommended. When this is done accurate and reproducible results are secured with no interference from nitrates and considerable latitude is permissible both in acid concn. and in time of hydrolysis. A. P. L.

Biochemical transformation of acetic acid by molds and the chemistry of citric acid formation. T. CHIRZASZCZ AND D. TILKOW. *Biochem Z* 229, 343-57(1930).—In the metabolism of AcOH by the mold *Penicillium succinace* fumaric, oxalic, l-malic and citric acids are formed and probably also other acids not yet identified. The finding of l-malic and citric acids is here reported for the first time. It is concluded that the citric acid formation in sugar cultures also proceeds over AcOH. The AcOH is changed to succinic acid and by dehydrogenation, and this in turn is dehydrogenated to fumaric acid. The fumaric acid is changed to l-malic acid by taking up a mol of H_2O . It is assumed that at some moment both malic and acetic acid ions are present in the culture, citric acid being formed from these two by a dehydrogenation. The transformation of the sugar to citric acid proceeds over the intermediate stages of pyruvic acid \rightarrow acetaldehyde \rightarrow AcOH, then as outlined before. K and Na ions have a greater effect than Ca ions on the transformation of AcOH and the formation of citric acid proceeds better, in fact, in the presence of Ca ions alone only succinic and l-malic acids have been found. S. MORGULIS.

Carbohydrate metabolism of higher plants. Formation of methylglyoxal by enzymes of germinating seeds. CARL NEUBERG AND MARIA KOFEL. *Biochem Z* 229, 433-42(1930).—Peas or beans were sterilized by submersion for 40 min. in 0.1% $HgCl_2$, then washed thoroughly with sterile H_2O and germinated in Petri dishes. After 1-2 days the fresh sprouts were removed from the seed membranes and ground finely, and either extracted or made into dry preps. with alc. ether. Extracts were made from the latter by 17-20 hrs. digestion with H_2O contg. 3% toluene. The methylglyoxal was produced from added hexose-phosphate ester, whereby the enzyme was completely removed from the fresh material or from the alc. ether dry prep., and could not be found in the residue. Obviously methylglyoxal is the intermediate product in the desmolytic changes in the sugar mol. both in plants and animals. S. MORGULIS.

Microchemical study of hemicellulose in some plant cells, with special reference to its distribution in the protoplasm. R. C. MALHOTRA. *J Biochem (Japan)* 12, 341-9(1930).—Recent studies on hemicellulose seem to suggest that this substance is not merely present in the cell walls impregnated in cellulose, pectin or lignin, but actually is present as a reserve carbohydrate. Microscopic studies reveal that the hemicellulose is present not only in the cell wall but also as protoplasmic inclusions. Hemicellulose tests on tomato cells grown in warm and cold temp. seem to show that there is more hemicellulose in cells of plants grown in the cold than in the warm temp. S. MORGULIS.

Glucides and glucidic compounds of the brown algae. H. COLIN AND P. RICARD. *Compt rend* 190, 1514-6(1930).—Many algae have been analyzed for mannitol, laminarin and alginate. The figures are given. T. H. RIDER.

The relationship between intracellular Ca^{++} and cytoplasmic sexualization of horsetail spores. PH. JOYE-LAVERGNE. *Compt rend* 191, 865-6(1930).—The spores of *Equisetum arvense*, *E. limosum* and *E. maximum* are divided into 2 groups. There is a relationship between the sexualization of the groups and their intracellular oxidation-reduction potentials. T. H. RIDER.

The presence of tannoids in flowers. ST. IONESCO. *Compt rend* 191, 867-8(1930).—Flowers or their parts are boiled in 96% alc. or in $(CH_3)_2CO$ until decolorized, then soaked in distd. H_2O until soft again, when the presence of tannoids can be demonstrated by the production of color by treatment with $Fe_2(SO_4)_3$, U acetate, $Cu(OAc)_2$ or K_2CrO_4 . All parts of many flowers were found to contain tannoids. T. H. R.

The soluble enzymes secreted by Hymenomycetes. The quinones and the antioxidant function. L. LUTZ. *Compt rend* 191, 880-2(1930); cf. C. A. 24, 5796.—Quinone is shown to have an antioxidant effect, retarding the action of the soluble enzymes of Hymenomycetes. T. H. RIDER.

Callusing of cotton stem cuttings. H. E. REA. *Plant Physiology* 5, 575-85(1930); cf. C. A. 13, 139.—The treatment of cotton cuttings with 2.0, 1.0, 0.5 and

0.25% of aq. solns. of KMnO_4 for 20 min., prior to placing in propagators, reduced the percentage of callusing 37%. WALTER THOMAS

Preliminary study of elongation of roots of Georgia collards as affected by sodium luminal. T. W. PRATT. *Plant Physiology* 5, 587-97 (1930).—Pronounced retardation of root elongation by Georgia collards (*Brassica oleracea*) was produced by sodium luminal, used in concns. from 0.1 to 1.0%, within 1.5 hrs. after administration. The effect of the sodium luminal on the root elongation was immediate, the greatest retardation occurring during the first of the six 15-min. intervals of observation. Distilled water partially removed the depressing action of the sodium luminal. The stronger the concn. of the drug, the greater was the retardation of root growth. WALTER THOMAS

Effect of various methods of storage on the chlorophyll content of leaves. PHILIP A. HARRISMAN. *Plant Physiology* 5, 539-606 (1930).—Soy bean and nasturtium leaves dried at 98° suffered a considerable loss of chlorophyll (approx. 70%). Leaves dried at room temps. (16° to 24°) lost 20 to 30% of their chlorophyll. Even at the optimum temp. range for oven drying (45° to 60°) some chlorophyll may be lost. Low temps. are ineffective in preventing the loss of chlorophyll unless the leaves are actually frozen. Leaves frozen with "Dry Ice" retain all of their chlorophyll, also leaves desiccated at reduced pressure (7 cm. mercury) over concd. sulfuric acid at room temp. lost little or no chlorophyll. Leaves desiccated at reduced pressure in the absence of O_2 and CO_2 lost more chlorophyll than leaves desiccated in air at the same pressure. Conclusion.—Freezing is the most satisfactory method of storing leaves for later detn. of chlorophyll. WALTER THOMAS

Production of kojic acid by *Aspergillus flavus*. O. E. MAY, A. J. MOTER, P. A. WELLS AND H. T. HERRICK. *J. Am. Chem. Soc.* 53, 774-82 (1931).—With 20% dextrose soln., *Aspergillus flavus*, under favorable conditions, is capable of transforming more than 45% of the dextrose present and 55% of that consumed into kojic acid in 12 days from the time of inoculation from spores. $(\text{NH}_4)_2\text{SO}_4$ was unsatisfactory as nutrient N, NH_4NO_3 was most satisfactory, NaNO_3 was much less satisfactory. Better results were obtained with a KH_2PO_4 1 g. and $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$ 2 g. per l. than with KCl 0.1, H_3PO_4 0.054 and $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$ 0.5 g. per l. Arange from 30° to 35° promoted rapid vegetative development with high yields of acid. In a range of sugar concns. from 15 to 35% growth and acid production were satisfactory, the highest yield occurring at a concn. of 20%. The max. quantities of acid were most economically produced in 12-day cultures when the ratio of the surface area of mycelium to the vol. of the soln. had a value of from 0.3 to 0.5. C. J. WEST

Chemical changes in infectious chlorosis in leaves of Abutilon. H. V. EULER, W. HERTZSCH, S. MYRBAEK, D. ROYENHELM AND A. FORSBERG. *Arkiv Kemi Mineral Geol.* 10B, No. 13, 6 pp. (1930), of C. A. 24, 5332, 25, 131.—Those portions of the leaves affected with infectious chlorosis have much lower catalase activity, higher proportion of amino N, less chlorophyll, xanthophyll and carotene, and less tryptophan than the normal green portions. J. J. WILLMAN

Studies on the ripening of rice grains. T. TADOKORO AND M. ABE. *J. Faculty Agr. Hokkaido Imp. Univ.* 27, 343-87 (1930), of C. A. 22, 967, 2396.—Rice was harvested at 4 weekly intervals, the last sample being dead ripe. The ash and P decreased at first. The total N showed no definite tendency to change, but the ratio of protein to non protein N increased constantly. Of the 4 groups of rice proteins detd., the water-sol. group gradually decreased, while the NaCl-sol. and alc.-sol. groups increased. The alkali-sol. group was inconstant in its changes. All of these changes are just the opposite of those during germination. While the non protein N is decreasing, the amino N of the proteins is also decreasing, indicating a formation of higher proteins from lower. Oryzenin, the alkali-sol. protein, decreased in ash and P during ripening, but increased in N, S and amino N. Its isoelec. point changed toward the alk. side, and its $[\alpha]$ increased. The amt. of amino N liberated from oryzenin by ultra violet light increased with ripeness. Arginine increased at first and then decreased, while histidine increased continuously. Lysine decreased at the later stage, while cystine decreased and then increased. Thus the authors think that oryzenin is at its highest degree of condensation sometime before dead ripening. The ash and P content of rice starch decreased gradually while the sapon. value increased. The amt. of glucose formed by hydrolysis with HCl attained a max. before complete ripeness, and at this same period the $[\alpha]$ of the acetylated starch reached a max. and the content of acetyl its min. J. J. W.

A limnological significance of free H_2CO_3 (LINDERMAN) 14.

E—NUTRITION

FRIEDRICH B. STAECKE

Urinary syndrome of metabolic disturbances caused by lack of vitamin B in the rat. J. ROCHON *Bull. soc. chim. biol.* 12, 342-55 (1930), cf. C. A. 23, 1166.—In rats receiving *ad libitum* the diet of Randon and Simonnet, the C/N ratio of the urine and feces is practically const. for a normal animal on a given diet, although the actual value varies according to the age of the rat. On a diet deficient in vitamin B the ratio increases (more rapidly as the point of death is approached). Ingestion of yeast temporarily reduces the ratio. The consequences of lack of vitamin B are manifested sooner on a diet rich in carbohydrate. Death is not due to inanition, which would cause a fall in the C/N ratio. B. C. A.

Fractional dialysis of urine. A. BORIVY *Compt. rend. soc. biol.* 101, 722-4 (1929).—A vegetable diet increases the quantity of easily dialyzable substances in urine, while a meat, and particularly a milk, diet has the reverse effect. B. C. A.

The effect of irradiated ergosterol (vitamin D) on the calcium-phosphorus metabolism of the chicken. I. J. KING AND G. I. HALL *Can. Chem. Met.* 15, 22 (1931).—One group of 4 month old chickens was fed an adequate basal diet contg. all the essential elements for growth except vitamin D. Another received the same diet plus 2 drops of irradiated ergosterol voln. daily. Their growth was normal. Group No. 3 was fed excessive amounts of irradiated ergosterol equiv. to 250-1500 times that given the normal birds. These birds lost wt. rapidly, became strikingly anemic and weak-legged. In the rachitic birds there was an increased phosphatase content of the bones accompanied by a hypocalcemia of the serum and in the hypervitaminous birds a subnormal phosphatase content accompanied by a hypercalcemia. W. H. B.

Effect of acidic and basic diets in chronic nephritis. D. M. LYON, D. M. DUNLAP AND C. P. STEWART *Edinburgh Med. J.* 38, 87-108 (1931).—High basic, high acid and high purine diets of similar protein content were investigated for their effect in chronic nephritis. The basic diet alone brought about clinical improvement as well as favorable variations in the blood and urine chemistry. RACIUS BROWN.

The group of "water-soluble vitamins" and their relationship. III. The scurvy of rats. WERNER KOLLATH *Arch. exp. Path. Pharmacol.* 153, 359-80 (1930); cf. C. A. 23, 132.—Rats can be given a disease histologically identical with guinea pig scurvy by a proper diet. The addn. of alk. hematin to such a diet causes 75% of young rats to develop beriberi; their bones, however, show the same scorbutic changes. It is therefore suggested that the absence of a common factor underlies both beriberi and scurvy, other factors of age, species, or diet detg. which disease will appear. H. B.

Methylglyoxal as an intermediary product of carbohydrate catabolism. FAUST SCHNEIDER AND LANST WIDMANN *Klin. Wochschr.* 9, 1916-7 (1930).—Previous workers have demonstrated that methylglyoxal is formed by liver, kidneys, muscle cells and red cells in the course of sugar metabolism. By removing the substance (as the 2,4 dinitrophenylhydrazones) S. and W. have shown that it is formed by human red cells, even after hemolysis, and also by white blood cells. As much as 70-75% of the total glucose added passes through the methylglyoxal stage. Diabetic blood cells show no abnormality in this respect. Material from abscesses did not form the substance, even though the bacteria involved were quite active, either in pure culture or when added to fresh leucocytes. H. EAGLE.

Biological significance of milk peroxidase: Preliminary report concerning a relation between vitamin B and peroxidase. TAYUKU ARAKAWA *Tohoku J. Exptl. Med.* 16, 118-22 (1930).—Pregnant rabbits on a diet contg. no vitamin B secrete milk contg. no peroxidase. With the addn. of vitamin B (beriberol) to the diet, peroxidase appears in the milk. H. EAGLE.

"Aperoxidatic" milk as a possibly early symptom of avitaminosis B. An experimental study. KAZUO SUZUKI AND TAYUKU ARAKAWA *Tohoku J. Exptl. Med.* 16, 228-31 (1930).—In rabbits secreting peroxidase free milk on a diet contg. no vitamin B, the peroxidase reaction of the blood leucocytes remains as strong as normal. H. B.

Studies on the quantity of the glomerular filtrate. I. Glomerular filtrate in beriberi. FUSAKICHI NAKAZAWA AND HIROYE KUSAKARI *Tohoku J. Exptl. Med.* 16, 321-8 (1930).—The glomerular filtrate of normal humans is 200-150 cc. per min., as estd. by creatinine output (if it is assumed to be a simple filtrate, with neither tubular secretion nor resorption). In beriberi it is much lower, and may fall to as little as 10 cc. per min. With the administration of vitamin B, this returns to its normal value. H. EAGLE.

The toxicology of the irradiated products of ergosterol. WOLFGANG HEUBNER.

Nacke Ges. Wiss. Göttingen Math.-physik. Klasse No. 2, 149-63 (1933)—The irradiated products of ergosterol were prepd. according to the method of Windaus, some of the samples being irradiated in ether, some in benzene and others in alc. In most cases the light from Hg was employed. Expts. on rabbits which were fed with the ergosterol products showed various pathological effects centered chiefly in the aorta and arteries. Death, following a loss in wt., occurred in all cases. Calcification foci were found in the aorta upon postmortem exam. Data are presented in 13 tables. RUSSELL C. EAS

The feeding of milk goats with a ration consisting of oat straw and nutritive material but poor in vitamin A, and its supplement. A. SCHÖNERT AND K. BERTRAM. *Fortschr. 1, No. 1, 70-87 (1932)*—During a 430-day expt. 3 milk goats were fed on a ration of oat straw, dried scraps and nutritive matter, but lacking in hay or green food. To insure milk production it was necessary to add a salt mixt. of CaO , H_3PO_4 , and NaCl . Before the addn. of this mixt. the milk yield dropped but rose to normal again upon the addn. Adding wheat germ, which contains a small amt. of vitamin A, produced a temporary increase in the milk yield. The effects of a lack of vitamin A were evident in 2 of the animals. One of these did not become pregnant and the other gave premature birth to 2 dead kids. The one goat that was given vitamin A daily in the form of 10 cc. of cod-liver oil went through an apparently normal pregnancy and gave birth to two dead kids and one living kid. The latter showed good development. After 3 months on the vitamin-A-deficient diet, the milk showed a deficiency in vitamin A. A long time was required to bring the milk vitamin content back to normal. RUSSELL C. EAS

Nutritional study of the white school children in five representative counties of Florida. CHESTER F. ARMBRIST, OTTO D. ARBOTT AND GEORGIA WESTOVER. *Fla. Agr. Expt. Sta., Bul. 215, 5-64 (1931)*—A survey of defects in white school children as correlated with diet, wt., age, etc. C. R. FILLERS

Effect of the use of mineral oil upon the absorption of vitamin A. JEROME L. ROWLEY. *J. Nutrition 3, 345-51 (1931)*—The effect of mineral oil on the absorption of vitamin A is dependent on the amt. of vitamin A in the diet. When the quantity of mineral oil given to rats is comparable to the human therapeutic dose and the vitamin A only enough to produce subnormal growth, the rats are unable to withstand the effects of the mineral oil. However, if the amt. of vitamin A is adequate, mineral oil can be given without noticeable effects. If sufficient vitamin is given the amt. of mineral oil could be increased 50% without producing a slackening of growth or preventing the cure of severe ophthalmia. If this laxative is employed, it should be accompanied by a generous amt. of fat-sol. vitamins. Endography. C. R. FILLERS

Preparation of a concentrated source of the heat-labile vitamin B, free from contamination with the heat-stable factor G. HERBERT M. EVANS AND SAMUEL LUKATEK. *J. Nutrition 3, 353-74 (1931)*—20% EtOH is an efficient extractive of both the heat-labile and heat-stable, water-sol. vitamins, the heat-stable factor is in low concn. in the outer layers of the rice kernel and decreases further in abundance in the inner layers, whereas the heat-labile factor is in high concn. in the rice outer layers and decreases only slightly in the inner layers. Clay activated by the 20% ext. furnishes an excellent source of vitamin B free from the heat-stable factor. When the syrup resulting from concn. of a 20% ext. is treated with 90% alc. so that the resulting concn. is 50%, most of the heat-stable factor is removed, nevertheless, large losses in the heat-labile factor are produced in this way. When the final concn. is 50% these losses are much reduced, though somewhat more heat-stable vitamin is present. In the case of the rice polish, so little of the heat-stable factor is initially present that fractionation with the 50% alc. yields an ext. almost free of the heat-stable factor and very rich in vitamin B. Finally, adsorption with clay is more selective for the heat-labile vitamin and thus the last traces of the heat-stable factor are left behind. Quant. removal of substance B from cereals by means of Ba(OH)_2 was not achieved, though feeding of these cereals gave a high dosage of substance B with assured absence of the heat-stable factor. Endography. C. R. FILLERS

Factors which determine renal weight. VII. Protein intake and age. ELMOR M. MACKAY AND LOIS LORENZO MACKAY. *J. Nutrition 3, 513-6 (1931)*, cf. C. A. 24, 74—An extended study of the linear relationship between renal wt. and protein intake proved that it is not influenced by age. The empirical linear formula, $\text{protein intake} = 0.0183 \text{ renal wt.} - 2.74$, expressed the relationship in normal male rats 70, 131, 220 and 410 days old when receiving adequate amts. of protein in the form of casein. VII. Protein intake and sex. *Ibid.* 505-44—A linear relationship between renal wt. and protein intake was demonstrated for female rats at all ages. When receiving adequate amts. of protein in the form of casein it may be expressed: $\text{protein intake} =$

4-5(1930)—Osteoporosis, a disease of horses where the bones, especially of the head, become porous and soft, was caused by diets deficient in Ca. The heavy clay soils of Central Porto Rico are definitely deficient in lime and the crops raised thereon are similarly deficient, thus giving rise to the disease in animals when the crops are used for fodder. The addn. of sea salt, bone meal and lime to the ration is recommended.

C. R. FELLERS

Nutritional discoveries in relation to livestock feeding practices. G. BONSTEDT. *Sci. Agr.* 11, 65-73(1930)—General.

C. R. FELLERS

Mineral feed problem with dairy cattle and swine. G. BONSTEDT. *Sci. Agr.* 11, 74-7(1930)—Tabular data are presented showing the most common mineral deficiency diseases in livestock and the remedy.

C. R. FELLERS

Antiscorbutic power of fruit and flour. T. SCHLOSSBERG. *Compt. rend. soc. biol.* 102, 1104(1930). *Physiol. Abstracts* 15, 101—Orange juice and lemon juice prevented the onset of scurvy in guinea pigs on a diet which otherwise caused it. Powder milk and yeast added to flour had no preventive action.

G. G.

Value of fats in the feed of dairy cows. ARNOLD BUSCHMANN. *Tierärzt. J.* 1, 120-78(1930)—The effect of adding vegetable oils or oil rich seed meals to the ration of dairy cows was studied. Up to a certain limit, varying from 0.4 to 1.0 kg. oil per day per 1000 kg. body wt., the fat content of the milk was not changed or was increased less than 0.1%. Feeding larger amts. slightly decreased the milk fat. The limit was lowest for rapeseed oil, highest for coconut, with sesame and linseed intermediate. The chem. and phys. constants of the milk fat were altered to the same extent as though approx. 15% of the particular oil in question had gone directly into the milk as so much butter fat.

L. E. GILSON

Vitamin D efficacy of hog feeds. VALENTIN HORN. *Tierärzt. J.* 1, 104-205(1930)—Healthy hogs made greater gains in wt. in proportion to the amt. of food eaten when "Hemom," a corn irradiated phytosterol prepn., was added to the ration, but not sufficient to show a profit over the added expense.

L. E. GILSON

Does the addition of potassium iodide to the diet affect the composition of the intestinal and fecal flora? A. SCHMIDT and J. ROEDERICHSEN. *Tierärzt. J.* 1, 206-18(1930)—This question has been raised by the widespread adoption of table salt contg. KI. Pigs were given 0.20 mg. KI per day in their ration and compared with rats not given KI. There was no apparent difference in the counts of the various bacteria and fungi. Tables of bacterial counts are given.

L. E. GILSON

Vitamin D preparations in hog feeds. VALENTIN HORN. *Tierärzt. J.* 1, 295-307(1930) cf. preceding abstr.—Vigantol Merck and Hemom when fed in equiv. quantities were equally efficacious in restoring to health and promoting the growth of groups of sick coughing rachitic young pigs.

L. E. GILSON

Tryptophan and histidine as "anabolites." GEORGES FONTÈS and LUCIEN TIMVOLLE. *Compt. rend.* 192, 63-5(1931) cf. *C. A.* 25, 1575—Simultaneous injection of tryptophan and histidine allows a living organism (dog) in equi. with an unchanging diet, to make very clear gains in wt., which are attained rapidly and maintained for a long time. These gains may be explained only by an improvement in the general assimilation processes, and especially those which have to do with N metabolism. The name anabolite as proposed for these amino acids, seems justified.

L. K.

Metabolism of lipides. The role of the pancreas endocrine in the phenomena of pulmonary fat catabolism. I. I. NITZESCU and G. BRUNETTO. *Bull. soc. chim. biol.* 12, 827-48 849-72(1931), cf. *C. A.* 24, 5811—From the content and rate of catabolism of fats in lung tissue and blood, *in vivo* and *in vitro*, N and B conclude that the destruction of fats in the tissues is dependent on a localized thermolabile factor which is activated by insulin which functions as a coenzyme or kinase. The localized factor is also diastatic controlling the catabolism of glycogen.

C. G. KING

Resistance of chickens to parasitism affected by vitamin A. JAMES E. ACKERT, MARIAN FISHER McLENNAN and NAOMI ZIMMERMAN CRAWFORD. *Am. J. Hyg.* 13, 325-56(1931)—The resistance of growing chickens to the intestinal round worm, *Ascaridia lineala*, was lowered when the fowls, 4-7 weeks of age, were kept on a diet deficient in vitamin A for a period of 5 weeks. The larger no. of worms remaining in the chickens is attributed to a weakened peristalsis. The source of the vitamin A in the control animals was cod liver oil. The criteria for judging whether this vitamin affected the resistance of the chickens to the parasites were the no. and the lengths of the worms remaining in the chickens at the close of the expt.

N. A. LANGR

Resistance of chicken to parasitism affected by vitamin B. JAMES E. ACKERT and L. O. NOLF. *Am. J. Hyg.* 13, 337-44(1931)—Preliminary tests were carried out to ascertain if the vitamin B complex is a factor in the resistance of chickens to the in-

testicular round worm, *Ascaridia lineata*. The source of vitamin B was baker's yeast. Groups of chickens given the same no. of embryonated eggs of the parasite had more worms when their diets lacked vitamin B than did similar groups fed adequate amounts of this vitamin. The larger no. of worms is attributed in part to partial paralysis of the intestine due to lack of vitamin, the greatly weakened peristalsis probably aiding the worms in remaining in their habitat. It is suggested that yeast may contain a factor which is favorable to the growth of worms since larger worms were found in those chickens whose diet contained yeast. N. A. LANG.

Variations in the chlorine content of the blood serum and urine of mice subjected to acute experimental scurvy. L. R. HANCOCK AND ANDREW MICHAUX. *Const. rend.* 191, 1978 (1930), cf. C. A. 24, 5335. W. B. LIVING.

The effect of fluorine on the calcium metabolism of albino rats and the composition of the bones. I. J. MCCART AND H. H. MERRITT. *J. Biol. Chem.* 90, 215-220 (1931). Fluorine, especially when consumed in the more soluble form of NaF at levels of 0.0331 and 0.0624% of P, may cause the deposition of an apparently abnormal constituent in the bones or an abnormal deposition of a non-Ca constituent, possibly a fluoride of a mineral other than Ca, as evidenced by an increase (av. 1.3%) in the % ash in the bones above the normal. There is also an interference with the deposition of Ca in the bones (av. decrease in Ca content 10%), and the P content of the bone ash was not significantly affected by NaF feeding but the Ca/P ratio was depressed. When either CaF_2 or NaF was fed at the above levels, growth was inhibited entirely aside from any effect on food consumption. The higher level of NaF only lowered the food consumption also. The Ca balance was unaffected by either CaF_2 or NaF up to the level of 0.0331% P in the ration but at a level of 0.0624% the % of Ca that would otherwise have been retained by the exptl. rats was lowered. The peculiar effect of F on the development of the teeth was again demonstrated, the characteristic abnormalities appearing when either the sol. NaF or the insol. CaF_2 was fed. Bibliography of 37 references. A. P. LACROIX.

The physiological effects of diets rich in egg white. HENRY L. PARSONS, *J. Biol. Chem.* 90, 351-372 (1931), cf. C. A. 24, 1842. Diets rich in egg white specially cause a nutritional disorder and death in young rats weaned from stock rations. As little as 6% of dried liver fat for 3 days before the beginning of the egg diet protects the rats from this disorder. The swiftness of onset of the first acute symptoms suggests that the effect is due to a direct injury from the egg white rather than to a nutritional deficiency in the usual sense of the term. In rats surviving the 1st few weeks on a ration containing 16% of dried egg white, either raw Chinese, or from raw fresh eggs, symptoms developed which strikingly resemble pellagra even though the ration contained 20% of potent dried yeast. These symptoms were prevented or cured by 20% of dried beef liver (but not by 10%), no matter how far advanced the dermal or nervous involvement. Lard and raw linseed oil were ineffective as preventive or curative agents. The symptoms noted were loss of wt., decreased activity, an awkward gait in walking, a humped back, emaciation, increased shedding of hair, soiled fur, paws and tail, closed eyelids, bloody urine, and progressive spasticity. The results suggest caution in the use of as large quantities of egg white in hospital feeding as has been the common practice in many hospitals. The introduction of egg white into the somewhat restricted diet of a child may also perhaps need to be safeguarded with protective foods in somewhat the same way as its introduction into the diets of the rats used in the expts. A. P. L.

Ketogenesis in relation to the food of swine. V. ZAGARI. *Arch. intern. physiol.* 33, 191-222 (1930). Swine excrete considerable amounts of "ketone bodies" when fed carbohydrate diets. These are increased considerably upon fat or protein diets. Fats produce a greater increase than proteins. The NH_3 excretion does not parallel that of the ketones but bears a relation to the protein ingested as well as to the "ketone bodies" excreted. In this phase of intermediary metabolism swine contrast with the herbivora. C. M. MCCAY.

The metabolism of nitrogen. VII. Comparative influence of carbohydrates and fats upon the utilization of nitrogen. G. PARQUER. *Arch. intern. physiol.* 33, 185-95 (1930).—A series of diets was designed containing casein, egg albumin and gluten. These proteins were fed with either starch or butter. The N balances were determined. P supports his earlier thesis that 10-10% of the N intake cannot be accounted for by fecal N plus urine N plus body weight increase. He finds no influence of fats or carbohydrates upon the "lost" N. C. M. MCCAY.

Experimental studies on the influence of diet on the formation of biliary and renal calculi. 11. Feeding on a diet deficient in fat-soluble vitamins. 2. The quantitative change in potassium, sodium, calcium, magnesium, total cholesterol and total

fatty acids in the blood. KAZUMASA USUKI *Japan J Gastroenterology*, 2, 226-30 (1930).—Normal values upon the sera of 82 young rabbits (50-120 days) were established as follows:

	Range (mg %)	Mean (mg %)
Ca	26.25—17.56	22.37
K	17.56—8.23	13.36
Mg	3.76—2.40	3.24
Na	605.00—260.00	492.00
Total cholesterol	80.35—47.56	60.5
Total fatty acids	251.12—160.00	19.0

These same constituents were determined upon 72 normal rabbits after they were placed upon a diet deficient in vitamin A. In cases that developed eye symptoms there was a marked increase in serum Ca. All the blood constituents increased in cases of calculi formation independently of the appearance of symptoms of avitaminosis. Ca and cholesterol increased the most. C. M. McCAY

Vitamins. Physicochemical constants of serum and plasma of animals fed on autoclaved meat with or without yeast. FERNANDO NEGRÍ *Biochim terap spec* 17, 445-60 (1930).—Dogs, fed exclusively on autoclaved meat, do not live longer than 2 months. The body temp decreases especially in the last week of life. The pH and the alkali reserve decrease, the refractive index of the serum decreases constantly. The sugar and N in the blood decrease; the latter increases in the urine, creatine in the blood and creatinine in the urine increase considerably. The addition of yeast restores the animal almost to normal, only so innocent hyperglycemia was observed. A. E. M.

Changes in the central nervous system in some avitaminoses. SILVIO BRAMBILLA, *Sperimentale* 84, 537-52 (1930).—Histological investigations of the central nervous system in scurvy of guinea pigs and in rickets of rats showed only slight degenerative changes, higher in scurvy than in rickets. No reason was found to suppose a unity of the neuropathological picture in different avitaminoses. A. E. MEYER

Sauerkraut as a source of vitamin C. B. A. LAVROV AND NATALIE YARUSOVA *Biochem Z* 229, 115-27 (1930).—Hay and oats which have been autoclaved invariably produce typical scurvy in guinea pigs receiving this as their basic diet. Addition of sauerkraut to such a basic diet inhibits the development of the disease process, 30 cc of the juice being required for a prophylactic dose. If it is assumed that the amt. of vitamin C present in the min. prophylactic dose of a definite product is a vitamin unit, 1 kg sauerkraut juice contains only 33 units, 1 kg sauerkraut 26 units while apple juice (cider?) will contain 330 to 660 units per kg. S. MORCULIS

Hypervitaminosis in chickens. EARL J. KING AND G. EDWARD HALL, *Biochem Z* 229, 315-22 (1930).—Under the influence of overdosage with irradiated ergosterol no calcareous deposits were observed in the organs of chickens, nevertheless the bones were weak and bent. There was a decreased phosphatase content in the bones and in the kidney. S. MORCULIS

Studies in experimental scurvy. X. Adrenaline content of the suprarenals and the amount of the adrenaline-like substance in the serum of guinea pigs fed on a vitamin-C-free diet. SHINSUKE OHATA *J Biochem (Japan)* 12, 419-27 (1930).—In the serum of scorbutic guinea pigs there is a definite diminution in the amt. of adrenaline like substance as well as a decrease in the adrenaline content of the suprarenal glands. Although the adrenaline content of the gland decreases also during inanition, this decrease is not as great as during scurvy. The change in the adrenaline like substances of the serum is the same in inanition and scurvy. The effect on a vitamin C-free diet is, therefore, not purely an inanition effect. S. MORCULIS

The relation of carotene to vitamin A. BASHIR AHMED *J Soc Chem Ind* 50, 12-4T (1931). cf C A 23, 1941, 24, 402, 5801.—Colored palm oils from different sources were tested colorimetrically for carotene and biologically for vitamin A. Tests were also made upon oils in which the carotene had been partly or largely destroyed by hydrogenation. The results of these tests indicate that the yellow coloring matter in the vegetable oils is carotene and that this substance shows vitamin A activity. Certain charcoals adsorb carotene, during which process it is oxidized and so made biologically inactive. In fish liver oils, the pigment can be adsorbed without destruction of vitamin A. N. M. NAVLOR

Lung calcification in healthy and tuberculous rabbits caused by massive doses of irradiated ergosterol. H. SIMONNET AND G. TANSSET *Compt rend.* 190, 1526-9 (1930). cf C A 24, 2499.—Irradiated ergosterol increases pulmonary calcification in

normal rabbits in a ratio of 1:20, while in tuberculous rabbits the increase was as high as 1:80. In chronic tuberculosis the evolution of tubercles is checked. I, H. R.

Biochemical investigations on rubrene. M. JAVITZ and (MIR) I. FRIEDMAN. *Compt. rend.* 191, 882-4 (1930). Rubrene (I), $C_{15}H_{12}$ is a crystalline orange-red compound rich in double bonds, and is decolorized by fixation of O_2 . It is in many ways analogous to carotene (II), $C_{40}H_{56}$. Albino rats fed on a diet deficient in vitamin A were fed I without improvement in the vitamin A deficiency symptoms. 75% of I by mouth is eliminated in the feces. The remainder is probably altered in the body since I cannot be detected in the organs. I and II are therefore not physiologically similar.

T. H. RINA

Further observations on the toxic effects of irradiated ergosterol. J. C. HAYES. *J. Pharmacol.* 40, 351-72 (1930). Cf. C. A. 25, 131. Bread or milk added to a synthetic diet did not protect rats against the toxic action of ole. Irradiated ergosterol which contained a high concentration of vitamin D. A diet modified to contain protein, carbohydrate, fat and ash in the proportions found in a bread and milk diet conferred no protection. A diet containing a smaller proportion of Ca and P had no protective action against irradiated ergosterol with either high or low concentration of vitamin D. A bread and milk diet had no protective action against units of irradiated ergosterol containing more than 70,000 antirachitic units.

R. RINA

Further biochemical studies on the antineuritic vitamin. A. THURMOND, S. J. JENSEN and M. RICHARDSON. *U. S. Pub. Health Repts.* 43, 3191-3200 (1930). The potency of a vitamin salt mixture prepared from brewer's yeast by the methods described by Smith (C. A. 24, 1104-9) and Schell (C. A. 23, 1721) was compared with that of a sample of the crystals prepared by Jansen (C. A. 21, 2156). 0.01 mg. S and S prepns. contain 0.0062 mg. N. They have the same curative value as 0.01 mg. J's crystals containing 0.0002 mg. N. There is not sufficient evidence to conclude that the same compound is involved in both preps. S and S purified their prep. by J's method, but no crystals were obtained. Rats were placed on a diet deficient in the antineuritic vitamin, until polyneuritis developed. The ration was then changed to one deficient in both vitamins, and the antineuritic concentrate administered intravenously (0.1 mg. per day). The rats recovered from paralysis but failed to grow. Death, due to nutritive failure, followed in 22-32 days.

W. GORDON RICK

Catalytic vitaminic substances and their action in beriberi. ALEXANDER MITCHELL. *Zymologie, Zym., (Chim. Col. e Luch.)* 5, 11-7 (1930).—The experiments prove that besides having a strong action against putrefactive bacteria, lactic ferments also cure polyneuritis in pigeons.

R. SAKOBU

Further fractionation of yeast nutritives and their relationship to vitamin H and Wildiers' "blo." ROGER J. WILLIAMS and FRANKLIN M. BRADWAY. *J. Am. Chem. Soc.* 53, 783-8 (1931).—Experiments with Wildiers' original yeast culture show that his "blo." is not readily adsorbed by fuller's earth and appears to be the same as the nutritive which stimulates the growth of Geibule Mayer yeast. No conclusive evidence of its multiple nature is available. The yeast nutritives studied by Miller and his associates are shown to be distinct from the "blo." of Wildiers. The Toronto workers should be credited with the discovery of a new series of yeast nutritives. Yeast No. 678 of the Am. Type Culture Collection and old process baker's yeast appear to have more complex requirements than Wildiers' yeast. Evidence is presented for the existence of a fourth distinct nutritive concerned in the growth stimulation of yeast No. 678. It seems reasonable to suspect that several of the nutritives for this yeast may be components of "vitamin H," since the antineuritic vitamin has previously been shown to affect its growth very strikingly.

C. J. WINT

Optical activity of cystine preparations used for animal experimentation. J. R. HAAO. *J. Am. Chem. Soc.* 53, 808-9 (1931).—Marston and Robertson (C. A. 21, 1070) have criticized studies of the use of cystine because the optical activity of the protein was not noted. Of 7 preps. examined, the value for α_D^{25} varied from -207.4° to -210.6° ; one had the value -185.9° . This variation cannot be explained on the basis of 8 percent but is partly due to the prolonged warming of certain samples with hot H_2O in order to remove the tyrosine present. It is highly probable that any one of the 7 samples is suitable for animal feeding unless it is assumed that inactive cystine is highly toxic.

C. J. WINT

Kalk-virgin (Kalkvirgin) 17. Effect of nitrogenous fertilizers on potatues (NORTH, et al.) 15.

CANTAROW, ABRAHAM: Calcium Metabolism and Calcium Therapy. Philadelphia, Lea & Febiger, 216 pp. \$2.50.

HARRIS, GEORGE A. Ja. Diet in Disease. Philadelphia P. Blakiston's Son & Co., Inc. 404 pp \$4 Reviewed in *J. Am Med Assoc* 96, 382, *Ann Internal Med* 4, 852(1931)

HOLCK, H. G. O.: Diet and Efficiency. A Five-Year Experiment on Man. Chicago Univ. of Chicago Press \$1

KEHLER, LYMAN F.: Eat and Keep Fit. Washington, D. C. (1322 Park Road) the author 302 pp \$3 Reviewed in *J. Am Pharm Assoc* 19, 1160(1930), *Am J. Public Health* 21, 221(1931)

SCHNEIDER, ARTHUR Der Vitamingehalt der deutschen Nahrungsmittel. Teil I. Obst und Gemüse. 2nd ed Heft 8 of "Die Volksernährung" Berlin J Springer 40 pp M 2 40

Irradiation products from ergosterol. Soc. des USINES CHIMIQUES RHÔNE-POULENC Brit 335,277, June 25, 1929 In order to obtain substances having intense antirachitic effect, unconverted ergosterol is sepd. from the product obtained by the irradiation of ergosterol, by crystg. from org. solvents in which the ergosterol is slightly sol. such as alc., acetone and EtOAc (the sepn. being continued until the sol. residue in alc. soln. shows a dextro-rotatory power greater than 25° for the yellow line of 1lg) The active product may be obtained by distn. of the solvent or may be taken up in a vegetable oil The irradiation may be effected in EtOH at a temp. of 75° while agitating the soln. with an inert gas, preferably with light having few or no rays below 2750 Å U and is preferably stopped somewhat before the max. production of antirachitic vitamin is reached

F—PHYSIOLOGY

E. K. MARSHALL, JR.

Skin. III. Sugar content of the skin under physiological and pathological conditions. C. URBACH AND G. SICHER *Arch. Dermatol. Syphilis* 157, 160-82(1929), cf. *C. A.* 23, 174, 893—Values of the sugar content of the skin and blood resp. are, man 47, 98, rabbit 117, 109, guinea pig 119, 117, rat 80, 83, dog 60, 92, mouse 53, 108 mg per 100 g (cc) After administration of sugar the skin sugar increases like the blood sugar and returns to normal values The skin sugar is diminished by insulin The effect of pathol. conditions of the skin is described B. C. A.

Water and phosphorus compounds of the nerve on degeneration. R. M. MAY *Compt. rend.* 190, 1150-1(1930)—During the first month the water content of the degenerated sciatic nerve of the rabbit increases to 14% above the normal and thereafter falls to the normal value The total P diminishes progressively during the first 2 months to a third of its normal value The bipn P decreases after 100 days to a tenth of its normal value, when histological examn. shows complete reabsorption of the lipin products of the nerve The alc. sol. and the protein P after an initial rise diminish to a third of the initial values while the water sol. P shows a steady increase up to 35% above the normal value after 6 months Also in *Bull. soc. chim. biol.* 12, 934-53(1930) B. C. A.

Urea and ammonia in normal gastric juice. D. SIMICI, R. VLADESCO AND M. POPESCU *Compt. rend. soc. biol.* 101, 199-202, 202-4(1929)—The gastric juice of fasting persons contains 0.014-0.042% of urea and 0.0048-0.0098% of NH₃, the values being scarcely affected by a test meal Injection of urea (5-10 g in 20 cc of water) raised the values to 0.06 and at least 0.03% resp. In pathol. conditions not affecting the kidneys these values are not markedly changed, in those affecting the kidneys they are increased to 0.3 and 0.11% resp. B. C. A.

Dextrose in normal urine. II. ROTH *Compt. rend. soc. biol.* 101, 218-20(1929)—The amt. of dextrose in normal urine varies between 0.02 and 0.30%, depending on the time of day After fermentation with yeast there remains in the urine 0.01-0.02% of reducing non-sugar substances. B. C. A.

Non-dialyzable fraction of urine. A. BOIVIN *Compt. rend. soc. biol.* 101, 724-6(1929)—The wt., C/N ratio, S content and P content of the non-dialyzable fraction of human urine depend on the technic of the dialysis B. C. A.

Oxygen consumption of the heart in varying, approximately isotonic, work. L. BAUER *Z. Biol.* 89, 513-22(1930)—By the method of Fismayer and Quincke (*C. A.* 23, 2504), it is shown that the amt. of O used for each contraction of the heart does not increase proportionately with increase in pressure, and therefore in work done, but bears a linear relationship to the initial vol. before contraction B. C. A.

The equilibrium between glycogen and lactic acid. WILDER D. BANCROFT AND GEORGE BANCROFT *J. Phys. Chem.* 35, 195-214(1931)—The rate of formation of

lactic acid during fatigue and the quantity of lactic acid found in fatigued muscle can be explained by assuming that enzymes catalyze its formation from glycogen. The reaction is reversible, but glycogen may be stabilized by adsorption on protein. The formation of acid would be slow during anorectic rest for this reason. The rapid action as a result of heat rigor or administering CHCl_3 is due to liberation of glycogen from the protein. Narcotics might act by freeing adsorbed glycogen from protein or by freeing enzymes from adsorption on the lipids or by a combination of the 2 reactions. The lactic acid found in the muscle is dextrorotary. One would expect inactive lactic acid to be formed by ordinary chemical processes. The adsorption of glycogen on protein could account for the conversion of lactic acid to glycogen during recovery of a muscle. Some failures to find glycogen adsorbed on protein or as a result of the action of enzymes on lactic acid are due to faulty methods of analysis. The formation of lactic acid from glycogen in cancer follows a path different from that followed in normal tissue.

T. T. BROWN

Recent advances in the physiology of digestion. IV. The latestiae. II. II. DUKES. *J. Am. Vet. Med. Assoc.* 31, 235-9 (1931), cf. *C. A.* 25, 993—A review.

RACHEL BROWN

Studies in blood diastase. Factors which cause variation in the amount of diastase in the blood. CHARLES REID and B. NARAYANA. *Quart. J. Exp. Physiol.* 20, 305-11 (1931)—The diastase in the blood was determined by the method of Fyle. Observations were made chiefly on dogs but also on humans, rabbits and cats. The blood diastase decreased after meals, injections of glucose, glycogen and insulin, but increased after anesthetics and injections of starch. Since pancreatectomy and ligation of the pancreatic duct caused little or no change in the amt. of circulating diastase the pancreas is not the main source of blood diastase. Since definite variations can be induced by injections of insulin and glycogen, it appears likely that the circulating diastase is not entirely a waste product on its way to excretion. Evidence is adduced that variations in blood diastase are probably due to its being taken up or given out by the liver cells according to the requirements of the body with respect to the glycogen glucose equil. probably in association with insulin.

RACHEL BROWN

Placental hormones. J. H. COLLIP. *Can. Med. Assoc. J.* 23, 631-3 (1930); cf. *C. A.* 24, 5812—A brief résumé of work already reported, in which C. stresses his view that the human placenta contains at least 3 distinct active principles, one a product of placental tissue functioning, and actively related to the condition of pregnancy.

A. T. CAMERON

Notes on the clinical use of certain placental extracts. A. D. CAMPBELL and J. H. COLLIP. *Can. Med. Assoc. J.* 23, 633-6 (1930), cf. *C. A.* 24, 2781—Observations on 135 cases of deranged ovarian function led to the following conclusions: Immenin does not alter normal menstrual cycles. Its effect in dysmenorrhea has been particularly encouraging, and it appears to correct certain types of amenorrhea of recent origin. It definitely lengthens the cycles in polymenorrhea. It relieves the menopausal symptoms of recent origin, but not in castrates. An anterior pituitary like principle from the placenta arrests certain forms of metrorrhagia.

A. T. CAMERON

The basal metabolism during harvesting. G. FARKAS, J. GELDRICH and A. SZAKALL. *Arbeitsphysiol.* 3, 468-76 (1930), cf. *C. A.* 23, 587, 24, 5352—The basal metabolism of 6 harvesters was determined by the Douglas bag method during varying lengths of time, the longest for a period of over 7 weeks. The basal metabolism at the time of the most strenuous work varied from 1.20 to 1.36 cal. per kg. per hr. and was about 10% higher than the mean values obtained during the periods of idleness after harvesting and threshing were finished.

T. M. CARPENTER

The protein catabolism at extraordinary altitudes. A contribution to the question of the limits of acclimatization at high altitudes. A. LOEWY. *Arbeitsphysiol.* 3, 506-604 (1930), cf. *C. A.* 25, 988—Ten urines obtained from 6 members of Dyhrenfurth's expedition to the Himalayas at altitudes between 5150 and 7300 m. and preserved in air-tight sterile bottles containing powder thymol were analyzed at Davos. Expressed in % of total N the constituents averaged: urea 75.4, ammonia 5.03, preformed creatinine 3.74, total creatinine 4.80, uric acid 2.22, amino acids 0.94 and undetd. N 12.5%. The urines were more nearly normal than those collected by Hopf (cf. *C. A.* 23, 5223) in the second winter Olympic. This cannot be explained by the inhalation of O_2 as only one sample was obtained under this condition and its composition agreed well with that from the same person without previous inhalation of O_2 . The approach to normality may be explained by the extensive acclimatization at the high altitudes. The acclimatization began at 2700 m. and the urines were collected 49 to 80 days later.

Meanwhile the subjects had successively gradually attained the high altitudes at which the urines were collected. T. M. CARPENTER

The growth and metabolism of mice. PH. L'HÉRITIER. *Ann. physiol. physicochim. biol.* 6, 1-31(1930)

The semi-permeable membrane of the cell. ROBERT CHAMBERS. *Ann. physiol. physicochim. biol.* 6, 233-9(1930)—A brief summary of the evidence furnished by microdissection for the existence of a true cellular membrane. H. EAGLE

The demonstration of bile acids in the blood. M. JANKY AND FR. STEINBERG. *Arch. exp. Path. Pharmacol.* 153, 244-56(1930)—By use of the Pettenkofer reaction, bile acids are shown to be occasionally less than 0.025 mg % in normal blood, disproving the theory that they regulate cardiac frequency. In jaundice, however, they may reach as high as 2-8 mg %.

Androkinin (male sex hormone) in male blood. S. LORWE, F. ROTHSCHILD, W. RAUTENBUSH AND H. E. VOSS. *Klin. Wochenschr.* 9, 1407(1930)—By extn. of saponin blood with ether, the active principle can be demonstrated in small quantities (1-2 mouse units per l.), by injection into a castrated mouse, and histological examn. of the glands of the seminal vesicle. H. EAGLE

Oxygen consumption following bodily work. DIETRICH JAHN. *Klin. Wochenschr.* 9, 1757-9(1930).

The significance of raw egg white for the deposition of glycogen in the liver. I. ABELIN. *Klin. Wochenschr.* 9, 1759-61(1930)—Egg albumin promotes the deposition of glycogen in the liver. H. EAGLE

Lactic acid metabolism in pregnancy and its relationship to hepatic and thyroid function. KARL JULIUS ANSELMINO AND FRIEDRICH HOFFMANN. *Klin. Wochenschr.* 9, 1768-70(1930), cf. *C. A.* 25, 731—In pregnancy there is an increased formation of lactic acid in muscles, both resting and active, due to increased thyroid function. Similarly increased lactic acid is observed in hyperthyroidism and in circulatory decompensation. H. EAGLE

The role of lactic acid in muscle fatigue. W. FLEISCHMANN AND F. SCHEMINSKY. *Klin. Wochenschr.* 9, 1773(1930)—By reversing the direction of the elec. stimulus, a fatigued muscle can be made to contract, despite its lactic acid content, indicating that muscle fatigue does not depend upon this factor. H. EAGLE

The anterior pituitary hormone. W. FALTA AND F. HÖGLER. *Klin. Wochenschr.* 9, 1807-12(1930).

Investigations in the role of the liver in the intermediary metabolism. KARL PASCHKE. *Klin. Wochenschr.* 9, 1917-9(1930)—In liver damage, as, e.g., in toxic jaundice, there is an abnormal increase in blood amino acids following the peroral administration of amino acids, suggesting that deamination is a function of the liver. H. EAGLE

Effect of temperature upon the height of contraction and time of maximal contraction of skeletal muscles. I. Frog muscles. KISUO SUGAI. *Tôhoku J. Exptl. Med.* 16, 1-24(1930)—The time required for maximal contraction decreases with increasing temp. (10-40°), the temp. coeff. is 2 between 10° and 25°, but only 1.2 between 25° and 40°. The degree of contraction is usually a min. at 22-27°, it increases at either higher or lower temps., occasionally it remains quite const., and in a third group decreases uniformly with temp. II. Muscles of white guinea pigs. *Ibid.* 54-76—Contrary to the results obtained with the muscles of poikilothermic animals, the extent of contraction of guinea pig striated muscle decreases with temp. (10-40°). The latent period and the time required for maximal contraction both show a similar decrease with temp., the temp. coeff. being 1.2 between 40° and 25° and 1.8 between 25° and 10°. At the same temps., the contractility gradually decreases, but the max. contraction time remains unaffected. H. EAGLE

Effect of temperature upon the contraction curve with several peaks. KISUO SUGAI. *Tôhoku J. Exptl. Med.* 16, 25-53(1930).

Studies in the internal secretion of the pancreas. XV. TAIZO KUMAGAI AND SATORU UKAI. *Tôhoku J. Exptl. Med.* 16, 157-63(1930)—Aq. exts. from various portions of pigeon pancreas were tested for their hypoglycemic effect upon rabbits. Those prep'd from regions relatively rich in alpha cells were found to be distinctly more active, contrary to the theory of beta-cell activity. H. EAGLE

Indispensability of the suprarenal glands in causing nicotine hyperglucemia in rabbits. TATSUAKIKO INABA AND KOGORO OIKAWA. *Tôhoku J. Exptl. Med.* 16, 169-77(1930)—Double splanchnicotomy does not affect nicotine hyperglucemia in rabbits, but removal of the suprarenals prevents it completely. H. E.

The vascular effect of tissue fluids. SABURO FUKAMI. *Tôhoku J. Exptl. Med.* 16,

181 8(1930)—Perfusion of Ringer fluid through surviving tissue (rabbit leg) endows the fluid with vasodilating properties, not observed if H_2O or NaCl soln is used. The vasodilating activity is most marked $1\frac{1}{2}$ hrs after the fluid is prep'd, and gradually disappears in the course of 24 hrs. H. YAGAN

The resorption of parenterally introduced emulsified fat. SRINOSUKA Hotta *Tohoku J. Exptl. Med.* 16, 311-20(1930)—The number of very fine granules seen in blood plasma (at lymph ts) is an accurate measure of parenteral fat absorption paralleling the results obtained on chem. analysis. Absorption from serous cavities is very rapid, and is conditioned by the circulation of the underlying tissues. H. R.

Biochemistry and biophysics of the developing hen egg. 1. Influence of humidity. ALEXIS I. ROMANOFF. Cornell Univ. Agr. Expt. Sta., *Mem.* 132, 3-27(1930).—From the exptl. data it is concluded that the extreme conditions in the incubator, 80 and 40% relative humidity, have both direct and indirect influence on the developing chick embryo, as follows: The growth of the embryo was rather irregular, being somewhat hastened at high and retarded at low humidity. The cycles in the growth of the embryo were decidedly disturbed. The Ca metabolism in certain stages of the embryonic development was much better at high than at low humidity. The mortality of the embryo was noticeably increased during the 2nd critical period, at the 10th day of incubation, particularly at high humidity. The physicochem. constitution of the yolk sac was slightly changed; this might have affected the embryonic metabolism, leading to the low vitality of the embryo and to its susceptibility to environment. The water content of the shell membranes and the contents of the allantoic sac corresponded in the humidity condition, this might have made an obstruction to hatching through the restriction of free movement of the embryo and through the hindrance to pulmonary embryonic respiration. Bibliography. C. R. MILLER

Effect of the estrus-producing hormone on the growth of the mammary gland. C. W. TURNER AND A. H. FRANK. Mn. Agr. Expt. Sta., *Research Bull.* 145, 5-60(1930).—In the normal rabbit after continued estrus, the mammary glands show extreme extension of the duct systems resembling the naked branches of a tree. If pregnancy or even pseudo-pregnancy now ensues, the ducts develop lobules containing large nos. of alveoli, resembling the budding of leaves from the smaller branches. The daily injection of 20 rat units of the estrus-producing hormone recovered from pregnant-cow urine, for 30 days in male castrate rabbits and in female rabbits castrated previous to puberty caused the growth of the duct system of the glands equal to that produced during continued estrus in the normal female. The results obtained seem to warrant the conclusion, that the estrus-producing hormone will cause the growth of the ducts equal to that produced during estrus, that the rate of development is not hastened by increasing the dosage and that there was present in none of the oils used a hormone or hormones which would produce the type of growth characteristic of pregnancy. C. R. MILLER

Basal metabolism of young women. HUGHNA MCKAY. Ohio Agr. Expt. Sta., *Bull.* 465, 3-37(1930).—The Benedict-Roth metabolism app. was used in study the basal metabolism on 91 girls aged from 11 to 18 yr. O consumption averaged 197 cc per min; basal heat production 1364 cal, and av. basal cal. production was 25.4 per kg. Heat production per kg. decreased fairly regularly as age increased. No significant differences among the age groups were noted on O consumption. When age was disregarded, heat production per kg. decreased decidedly with increased wt. The overweight group averaged 9% less and the underwt. group, 12% more per kg. than the entire group. Predictions of basal metabolism on surface area are less liable to error than those based on height or wt. Basal metabolism of young college women is somewhat higher in the spring than in autumn or winter. C. R. MILLER

The effect of manganese on the growth of rats. V. E. NELSON, J. M. EDWARD AND W. L. SWARTZ. *Proc. Iowa Acad. Sci.* 36, 267(1929).—Mn in small quantities stimulated the growth of rats. The basal diet employed consisted of casein 18, yeast 12, cod liver oil 5, salts (mixt. no. 185) 37 and dextrin 613. $MnSO_4 \cdot 4H_2O$ was fed at 2 levels, 100 and 600 p. p. m. of ration. In 50 days the control gained 178 g. daily and consumed 431 g. of feed per 100 of gain. On the smaller Mn allowance the gain was 206 g., and the feed requirement was 391 g. for 100 gain. On the larger Mn allowance the figures were 169 and 158 g. Apparently the 100 parts of $MnSO_4 \cdot 4H_2O$ per million inhibited growth. W. G. GAUSMUR

Hydrolysis of levorotatory substance in human milk. M. YOLONOVSKI AND A. LEPAGNOL. *Compt. rend. soc. biol.* 101, 61-6(1929); *Physiol. Abstracts* 14, 373.—A levorotatory non-reducing lactoside, a glucoprotein, exists in human milk, which is converted to a reducing substance by hydrolysis. G. G.

Basal metabolism as a function of temperature. C KAPER *Compt rend soc biol* 101, 708-10(1929), *Physiol Abstracts* 14, 653—The influence of the surrounding temp on the min heat production of pigeons after a 26 hrs' fast has been detd. By a simple mechanism of adaptation the exchanges in a pigeon may be diminished by 20-45%. The adaptation is slow, and may take several weeks for completion. In contrast with this slow mechanism is the rapid chem regulation setting up increased heat production when the temp is lowered. G G

Coagulation of blood and cholesterol content. L BUGNARD *Compt rend soc biol*, 102, 231-3(1929), *Physiol Abstracts* 14, 552—The cholesterol of serum is greater than that of plasma. The coagulation of blood under o₂ or in air is accompanied by an increase of cholesterol in the fluid phase. G G

Mechanism of concentration of urinary indoxyl. G LAROCHE, A GRIGAUT AND P DELILLE *Compt rend soc biol*, 102, 532-4(1929), *Physiol Abstracts* 15, 38—Parenteral injection of isotonic saline soln in the dog leads to a lowering of urinary indoxyl. Ingestion of water by the mouth increases it because of the absorption of indole from the alimentary canal. Peptone shock causes an augmented output. Tissue disintegration with transformation of tryptophan to indole may account for some of the increase. This would imply the endogenous formation of indole, which cannot be demonstrated. G G

Cholesterol regulation in the lung. L BUGNARD *Compt rend soc biol* 102, 550-1(1929), *Physiol Abstracts* 15, 18, cf C A. 24, 647; 4816—The distribution of cholesterol in the blood depends on the reaction. When the acidity is increased the amt of cholesterol in the plasma rises, the effect being explained by a transference of cholesterol from the corpuscles. Diminution of CO₂ in the lung is accompanied by the reverse change in the cholesterol distribution. The distribution seems to be a factor in maintaining the physicochemical equil and the p_H of the blood. G G

Oxidation-reduction potential of the vitreous humor. E REDSLOB AND P REISS *Compt rend soc biol* 102, 1060-2(1930), *Arch phys biol* 7, 221-5(1930), *Physiol Abstracts* 15, 145—The oxidation reduction potential of the vitreous humor in the rabbit was measured by the injection of indicators and by the insertion of a fine platinum Ir point with an indifferent electrode on the skin for elec measurement. The results of the 2 methods agreed well. A potential of +110 mv was found, which corresponds with an *r_H* of 10.1 for a p_H of 7.6. Arrest of the circulation caused a fall of potential to +81 mv, corresponding to an *r_H* of 16.9. G G

Reactions of ovary to placental injections. M KLEIN *Compt rend soc biol* 102, 1068-9(1930), *Physiol Abstracts* 15, 181—Exts of total placenta cause a great disturbance of the ovary in the rabbit. The changes resemble those obtained with hypophysis. The effects are produced with human placenta at all stages of pregnancy and at term. G G

Is the active placental substance a hypophyseal hormone? M KLEIN *Compt rend soc biol* 102, 1070-1(1930), *Physiol Abstracts* 15, 181-2—Hemorrhagic follicles and corpora lutea are found in the ovary after injection of the placental ext. These changes are supposed by some workers to be sp tests of hypophysis. The increased thyroid activity described as resulting from injections of hypophysis did not occur as a result of placental injection. It is probable that the placenta contains a substance which acts on the ovary but which is not hypophysis. G G

Acidosis in pregnancy. E L SOLAI, J DALSACE AND M GIRARDOT *Compt rend soc biol* 103, 111 2(1930), *Physiol Abstracts* 15, 184—Uncompensated acidosis has been demonstrated in albumine pregnancy. Now it is shown that acidosis exists to some extent toward the end of normal pregnancy. It is greatest in labor and disappears after parturition. There is a lowering of p_H and of the alk reserve. In pre-eclamptic conditions the p_H and alk reserve are lowest. G G

Oxytocic properties of woman's blood during parturition. J FONTES *Compt rend soc biol* 103, 349-50(1930), *Physiol Abstracts* 15, 119—The oxytocic substance in woman's blood must be a fetal product or a product obtained from fetal accessory structures, as it disappears a few hrs after parturition. The blood of the umbilical cord is not active, while that from the placental site of the uterus is. The placenta may be the seat of formation of the oxytocic substance. The oxytocic element is the detg factor in setting up labor pangs. G G

Secretory phenomena in the thyroid of birds. P FLORENTIN AND M WEIS *Compt rend soc biol* 103, 501-3(1930), *Physiol Abstracts* 15, 179-80—In certain birds massive degeneration of the parenchyma of the thyroid was observed. The debris is poured directly into the circulation, cellular emboli being seen in the small vessels. Amitotic multiplication is supposed to be the method of regeneration. The

analogy with mammalian thyroid processes is marked. The phenomenon seems to be connected with changes in the sexual organs. G. G.

Difference in cerebral effects of a new thyroid extract and thyroxine. G. J. LING. *Compt rend soc biol* 103, 641-5 (1930). *Physiol Abstracts* 15, 180.—An ext. has been made from the thyroid of vagotome animals which acts on psychomotor centers in lowering their chronaxie. It contains a hormone which is different from that affecting metabolism. Thyroxine does not lower the chronaxie of the nervous centers. G. G.

Substances extracted from the suprarenal glands by hydrogen peroxide. J. STIER. *Compt rend soc biol* 103, 650-2 (1930). *Physiol Abstracts* 15, 180.—In extg. the cortical hormone of the suprarenal gland, adrenaline may be excluded by using H_2O_2 . The ext. contains no lipoids, but choline is abundant. The alc. ext. contains lipoids and a crystallizable substance. The crystallizable substance is supposed to be identical with the oxidation product of adrenaline. Choline exists only in the peroxide ext., not in the alc. It probably arises from the decompos. of lecithin. G. G.

Remarks on the work of Mond and Netter "does the ionic permeability of muscle change during muscular activity?" I. LASSER. *Arch ges Physiol (Pflügers)* 226, 244 (1930), cf. C. A. 24, 5354. Polemical. ARTHUR GRUHLMAN.

Hematin, bilirubin and porphyrins. HANS FISCHER. *Naturwissenschaften* 18, 1020-7 (1930), cf. C. A. 25, 520.—An elaborate review of the present knowledge of the chemistry of blood pigment and related chem. compds., numerous references are given. B. J. C. VAN DER HOUTEN.

Morphology of the porphyrins. M. BOKST. *Naturwissenschaften* 18, 1015-41 (1930).—A review of recent work on morphological distribution of porphyrins and derivatives (cf. Birst and Konigsdorff, *Untersuchungen über Porphyrine*, Leipzig, 1929). B. J. C. VAN DER HOUTEN.

Proteoclastic function of the thyroid. M. LÖFFLER, A. LEMARE AND J. TUNST. *Compt rend soc biol* 102, 279-80 (1929). *Physiol Abstracts* 14, 589.—The thyroid gland breaks up proteins. Nucleoalbumins (precipitable by $AcOH$) are less abundant in the venous than in the arterial blood. Serine is more abundant in the venous blood. The amino acids are less abundant in the efferent blood. Fresh thyroid ext. breaks up albumins *in vitro*. B. C. K.

Resi-nitrogen percentage of the blood of natives and of Europeans in the Tropics. W. RADNAG. *Mededel. Dienst Volksgezondheid Nederland-Indië* 19, 1^e 2, 227-32 (1930).—A series of urea exstns. with undernourished natives resulted in an av. of 21 mg %; from 3 to 4 hrs. after the last meal at 10 or 11 A. M., the av. was 25.8 mg %.

The blood exampl. was capillary blood from the fingertip. J. A. KENNEDY.

Gasometric determination of the relative affinity constant of carbon monoxide and oxygen in whole blood at 38°. J. I. SANDROV, S. H. LIU AND D. D. VAN SLAER. *Bull. soc chim biol* 12, 512 (1930).—Three normal and 3 pathol. samples of human blood gave a value of 210 for the affinity const. K ($\mu\text{HCO}_2/\mu\text{HCO}$) = $(\mu\text{HCO}_2/\mu\text{HCO})$, and for 10 samples of beef blood $K = 170$. All values found were within 2.5% of the av. C. G. KING.

Food balance and nutrition. LUCIEN RANDUIN. *Bull. soc chim biol* 12, 815-20 (1930), cf. C. A. 24, 5805.—Pigeons receiving 64% carbohydrate on a vitamin B free diet died in 17-27 days, but when on a carbohydrate free diet (54% protein) they survived 3 months. Similarly rats died in 30 to 40 days or survived 4 months. Expts. also indicate a needed balance of mineral salts and carbohydrate. Details concerning purification of the diets are not given. C. G. KING.

Relationship between electrical differences of potential in the skin and normal basal metabolism. CHARLOTTE LUDV., A. FRANCES JOHNSON AND CHARLES SHILLARD. *Science* 73, 46-9 (1931). W. D. LANGELEY.

Regulation of cholesterolemia. L. BEGGARD AND C. SOUTA. *Compt rend* 191, 1382-4 (1930).—The regulation of cholesterolemia is an automatic function of the blood tissue. It is accomplished by the globular cholesterol existing in this tissue although it is neither generated nor destroyed by it. The mechanism of such regulation is purely a physical one and depends upon the cholesterol circulating generally in the cellular constituents of the blood and in the plasma. B. S. LIVING.

The work of the kidney in the production of urine. HENRY BORSOOK AND HOWARD M. WINIGARDEN. *Proc. Nat. Acad. Sci.* 17, 3-12 (1931).—Summary: (1) The work of the kidney in the excretion of urine is analyzed by means of the second law of thermodynamics. (2) It is shown that the work performed by the normal kidney in min. in the excretion of urine is of the order of magnitude of 0.7 g. cal. per cc. of urine, or 70 g. cal. per g. of N excreted. B. S. LIVING.

The energy cost of the excretion of urine. HENRY BORSOOK AND HOWARD M.

WINEGARDEN. *Proc. Nat. Acad. Sci.* 17, 13-25(1931)—Summary: (1) The work of the kidney in the excretion of urine is analyzed by means of the second law of thermodynamics. It is shown that the production of either hyper- or hypo-tonic urine entails work on the part of the kidney. The excretion of a urine which is the same as the plasma in all details incurs no work by the kidney. (2) The energy consumed by the kidney in man in the production of urine was found to be 6-11 kg. cal per g. of N excreted. (3) It is shown that the normal healthy kidney considered as a chem. machine possesses a great capacity for work, but performs its work with an "efficiency" probably not greater than 1-2%. (4) The effect of disease is to reduce markedly the capacity of the kidney for work. B. S. LEVINE

Gas tensions in the tissues. J. ARGILL CAMPBELL. *Physiol. Rev.* 11, 1-40 (1931)—General review covering diffusion, absorption, quant. measurement and other problems concerning gases in body tissues, cavities and secreted fluids. Extensive bibliography. E. R. LONG

Factor influence on oxidative and lytic systems. W. LOEHL. *Virchow's Arch. Path. Anat.* 279, 237-13(1930)—In the oxidation system, $\text{OH} + \alpha\text{-naphthol} + \text{NH}_2(\text{COOH}) + \text{COH} + \text{Fe}^{+++}$ a variety of amino acids may take the place of the glycocoll. This artificial oxidation system, in its susceptibility to heat, prolonged standing, ZnCl_2 , etc., behaves much like the cell oxidase supposed to function in the granules of eosinophilic leucocytes. Amino acids function in tryptic systems *in vivo*, inhibiting the action of alkali. While *in vitro* the groups COH and $\text{NH}_2(\text{COOH})$ both inhibit the action of alkali, they are themselves antagonistic, the inhibiting action of COH being removed by NH_2 . In the formation of blood cells peptic processes occur in the lymphocytes and tryptic processes in the granular leucocytes and red blood corpuscles. The formation of these cells may be modified by changing the OH concn. of tissue, and by treatment with aldehydes, amino acids or ferrous salts. E. R. LONG

The equilibrium between cerebrospinal fluid and blood plasma. III. The distribution of calcium and phosphorus between cerebrospinal fluid and blood serum. H. HOUSTON MERRITT AND WALTER BAKER. *J. Biol. Chem.* 90, 215-32(1931)—Normal serum contains from 9.35 to 10.6 mg. of Ca per 100 cc. (av. 10.0 mg.) while the Ca content of cerebrospinal fluid is approx. $\frac{1}{2}$ this amt. (4.5 to 5.23 mg. av. 5 mg.) The Ca content varies directly with the serum protein in such a way that there is no significant change in the ratio with changes in the serum protein. In meningitis there is a slight diminution in serum Ca and an increase in cerebrospinal fluid Ca which is accounted for by increased permeability of the meninges and the choroid plexus with a resulting increased protein content in the spinal fluid. There is a definite decrease in serum Ca at the end of pregnancy; the amniotic fluid Ca averages 6.59 mg. per 100 cc. In non-suppurative diseases of the central nervous system the Ca content of both serum and spinal fluid is normal; the spinal fluid P content varies between 31 and 45% of the serum P (av. 38%) in these diseases while in suppurative conditions the ratio varies between 35 and 123% (av. 40%). The serum and spinal fluid Ca content of cats is slightly higher than in humans, the spinal fluid Ca averaging 54% and the aqueous humor 63% of the serum Ca. IV. The calcium content of serum, cerebrospinal fluid and aqueous humor at different levels of parathyroid activity. *Ibid.* 233-45—The cerebrospinal fluid Ca remains markedly constant in spite of a well-marked change in the serum Ca level, the striking lowering of serum Ca in parathyroid tetany and its sustained elevation following parathormone administration being without appreciable effect on the spinal fluid Ca level. Injection of parathyroid hormone causes only a slight increase also in the Ca content of the aqueous humor (0.6 mg.). If one accepts the cerebrospinal fluid as a dialyzate in osmotic and hydrostatic equil. with the blood plasma, the data would indicate that variations in serum Ca occurring during parathyroid tetany or following the administration of parathyroid gland ext. are chiefly and perhaps wholly limited to the non-diffusible Ca. Variations in serum Ca associated with different levels of serum protein involve both the diffusible and non-diffusible Ca. Bibliography of 55 references. A. P. LOTHROP

Studies in hyperthermia. II. The acid-base equilibrium in hyperthermia induced by short radio waves. FRITZ BACKOFF, M. LOTISA LONG AND ELSIE HILL. *J. Biol. Chem.* 90, 321-9(1931) *cf. C. A.* 24, 2784.—"With the exception of a fall in blood vol., which is probably not a direct effect, no difference in effect was noted between raising the body temp. by placing the subject in the field of condenser plates in circuit with a short wave radio transmitter and in raising the body temp. by means of diathermy, warmed air or hot water baths. Of fundamental importance was the loss of CO_2 . The pH of the blood became more alk., there was a shift of bases to the blood proteins and an increased oxygenation of the hemoglobin of venous blood. No signifi-

ent change in the non protein N constituents of the blood was noted. No evidence was obtained that the body was attempting to compensate for the condition of alkalosis through the urinary or perspiratory excretions." III. The phosphorus equilibrium. FRITZ BISCHOFF, L. C. MAXWELL AND ELSIE HILL. *Ibid* 531-9 — No P is eliminated in the perspiration during a hyperthermia induced by rad waves. The decrease in the rate of urinary P excretion is accounted for by the decrease in P elimination with increase in metabolism and the conversion of inorg. P in the blood with an alk. shift of blood pH. The decreased rate of urinary N excretion is accounted for by the increased N elimination through the perspiration. A. P. LORIMER.

The action of general cold stimuli (sea bathing) on the resistance of erythrocytes. A. CHIAVELLINO. *Arch. sci. Biol. (Italy)* 15, 346-52 (1930). *C. A.* 24, 4824 — Cold stimuli such as sea bathing cause an increase in the no. of red blood corpuscles and a definite but slight decrease in the resistance of the corpuscles to hemolysis. P. M.

The place of origin of the biliary pigment. EMILIO BECCATI. *Arch. sci. Biol. (Italy)* 15, 356-65 (1930). — A review. PETER MASUCCI.

The reaction of articular tissues. BARTO BOSCHI. *Memoria med.* 1930, 11, 913-7 — The object was to see whether intensive work, immobility and the local action of heat and cold have the power to modify the pH of articulations. One knee of the rat was used for the study while the other was kept as a control. The pH was determined by means of phenol red which was injected intraperitoneally into the rats. The following results were obtained. Intensive work changed the tissues of articulation toward the alk. side. immobility did likewise at first, but later the reaction of the tissues changed toward the acid side of the normal. The local application of heat increased the alk. Cold produced slight and inconst. changes toward the acid side which was revealed by the diminution of the buffer value of the tissues toward acids. The mechanism of these changes and their significance in the general pathology of articulation are discussed. PETER MASUCCI.

The curve of alveolar carbon dioxide tension during voluntary apnea. GIORGIO GIORGI. *Arch. sci. Biol.* 84, 11-22 (1930). — For the same individual there is an average value for alveolar CO_2 tension beyond which it is impossible to maintain apnea. This max. alveolar CO_2 tension varies for different individuals, but for the same individual it is fairly const. S. MOROZZI.

Influence of the liver on the function of the heart. A. RONCATO. *Arch. sci. physiol.* 84, 23-31 (1930). — Blood perfused through the liver produces a marked hemodynamic effect upon the isolated heart. The perfusion arrangement is such as to permit the inclusion or exclusion of the pancreas, and this shows that the pancreas does not contribute to the effect upon the heart of the blood circulating through the liver. No conclusion is drawn as regards the nature of the substance given off by the liver which may be a hormone, or some nutritive factor furnished by the liver to the heart. Finally, it is also considered that the liver may detoxicate the blood which passes through it before reaching the heart and thus increase its action upon the contraction of the heart. S. MOROZZI.

Studies of carbohydrate metabolism of blood and organs with the aid of a new colorimetric method. I. The presence of different hexoses besides glucose in human blood and their formation in vitro. ZACHARIAS DISCHE. *Biochem. Z.* 229, 109-95 (1930). — Two methods were employed for the detn. of the blood sugar. A part of the blood is deproteinized with 25%, and another part with 10% $\text{CCl}_4\text{CO}_2\text{H}$. After remaining overnight this is centrifuged. It has been definitely established that no free sugar is retained in the ppt. nor do any substances pass into the fluid which affect the color reactions utilized in the sugar detn. Added glucose has always been quantitatively recovered. The color reactions were produced with diphenylamine-HCl or with indole- H_2SO_4 . The former reaction was so modified that it included all hexoses and not merely fructose. For this purpose to 1 part of the sugar soln. are added 2 parts of Rotherfusser's reagent (100 cc. concd. HCl, 80 cc. glacial AcOH, 20 cc. of 10% alc. diphenylamine); this is then heated in the boiling water bath 3, 10 or 30 min. The ketoses show a much stronger coloration than aldoses, the difference being greater the shorter the time of boiling. By the indole- H_2SO_4 reagent the relative color intensities for various sugars are as follows: glucose 100, fructose 125, Harden-Young ester 100, mannose 70 and galactose 75. The results by the former reaction are invariably higher (10-150%) than by the latter. In some instances the diphenylamine-HCl procedure gave variable results depending on whether the heating was for 10 or 30 min., which is regarded as an indication that blood contains 2 other hexoses besides glucose, an aldose and a ketose, which are designated as the "additional sugar." Following a large glucose dose the intensity of both the diphenylamine-HCl and indole- H_2SO_4 reaction gen-

erally increases to the same extent, so that only the glucose but not the "additional sugar" of the blood is raised under these circumstances, but occasionally there is an increase also in the "additional sugar". An hr after the injection of 20-30 units of insulin both glucose and the "additional sugar" disappear, but the latter disappears more rapidly. Incubating blood at 37° for 2 hrs causes practically all the "additional sugar" to disappear. From the behavior of fluoride or citrate blood kept for several hrs at room temp it is evident that a ketose sugar is formed, which disappears more quickly than glucose when the blood is kept in the incubator. Addn of glucose to citrated blood causes the formation of a ketose as well as an aldose. The newly formed ketose is present only in the red cells whereas the preformed "additional sugar" is practically equally distributed between cells and plasma. The formation of the ketose in blood at room temp is much increased through the addn of Na_2HPO_4 , and it is believed that this ketose is really a hexosephosphate ester. S MORCULIS

Influence on the development of the female sex organs of hyper-interfertilization. CARL MÜLLER. *Endokrinologie* 8, 5-14(1931)—Expts were performed on female rats exactly 3 weeks old. Some of these young rats were thymectomized, and only about 15% of the operated animals were lost. A special ext. of the adrenal bodies was used to inject some of the intact and of the thymectomized rats. For this purpose very fresh beef suprarenal glands were freed from fat, etc., and the cortex, carefully sepd from the medulla, was ground fine in a few cc. 0.9% NaCl. The thick, reddish pulp obtained from about 150 g. substance is boiled in an Erlenmeyer flask for 10 min on the water bath to remove proteins. This is filtered through gauze, about 30 cc. soln being obtained. Traces of adrenaline are removed by passing air through the ext. which is now neutralized with 0.05 N NaOH. Every 2-3 days 0.5 cc. of this was injected. These injections of the cortical ext. produced a typical strong inhibition of development of the external as well as internal female sex organs. In the ovary defective development of the ova was observed and absence of normal corpora lutea. S MORCULIS

Inactivation by proteolytic enzymes of the anterior hypophysis hormone obtained from pregnant urine. M REISS, A SCHÄFFNER AND F HAUROWITZ. *Endokrinologie* 8, 22-4(1931)—Expts were made with the following proteolytic enzymes: aminopolypeptidase, carboxypolypeptidase, pepsin, inactive trypsin and activated trypsin. Only the activated trypsin brought about complete destruction of the hormone. S MORCULIS

The significance of the bile acids in the carbohydrate metabolism. IX. The effect of bile acids on the sugar assimilation. Sei FUJITA. *J Biochem (Japan)* 12, 383-91(1930)—The glycogen forming capacity of the fasting rabbit liver from different monoglucides decreases in the following order: d fructose, d glucose, d mannose and d galactose. The glycogen forming capacity in this series can be raised on the av. 42-52% by the administration of cholic acid. Fructose yields the greatest glycogen production both with or without cholic acid, in point of extent of the glycogen formation with cholic acid this is followed by glucose, mannose and finally galactose. The lowering of the excretion threshold for these various sugars under the influence of cholic acid is therefore to be attributed to its stimulation of glycogen formation in the liver. S MORCULIS

Studies on the nature of the porphyrin appearing in urine following ingestion of chlorophyll. YASUO NAKAMURA. *J Biochem (Japan)* 12, 475-85(1930)—Spectroscopic studies as well as biol. tests show that urine porphyrin following the ingestion of chlorophyll is isomeric with blood porphyrin. The urobilin output after chlorophyll ingestion is always increased. S MORCULIS

The site of ammonia production. A IGILLING. *Skand Arch Physiol* 61, 23-6(1931)—I experimented upon myself on an ordinary diet, and in a condition of acidosis caused by NH_4Cl or alkalosis by NaHCO_3 . The plasma NH_3 was detd by the Parnas method, the rate of kidney filtration by Rehberg's method and the urinary NH_3 in hourly samples. So long as the urine is more acid than pH 7.8 more NH_3 is found in the urine than is filtered through the kidney during the same time, so that the NH_3 must be formed by the kidney. The results of analysis also allow the calcn. of the concn. of undissocd NH_4OH in plasma and urine, according to which the latter is 10 times as great as the former, thus precluding any possibility of a direct diffusion from plasma into urine. S MORCULIS

Further observations on the depressor substance in brain tissue. RALPH H MAJOR AND C J WEBER. *J Pharmacol* 40, 247-52(1930), cf C A 24, 2109—The depressor substance is adsorbed on Lloyd's reagent from a brain ext. and then freed by treatment with Ba(OH)_2 . The soln. is further purified by pptn. with phosphotungstic

acid and AgNO_3 . The final soln produces a fall in blood pressure in dogs before or after atropinization. It gives a positive Sakaguchi, and a negative Pauly, reaction. The substance is not histamine, choline or acetylcholine. C RUGEL

G—PATHOLOGY

H GIBSON WELLS

Chloride determinations in cases of intestinal obstruction. M PAGET *Bull soc chim biol* 12, 407-10(1930) — Chloride determination in cases of high or low obstruction of the intestine should be made on whole blood when the alk. reserve is increased and on the serum when it is normal. B C A

Magnesium and calcium in the liver of hyperthyroidized animals. M CAHANE *Compt rend soc biol* 100, 577-8(1929) — With guinea pigs the increase was 0.0021 and 0.0015%, resp. B C A

Relation between changes in the alkali reserve and relative chlorine and sodium contents of the blood in experimental immobilization of the intestine and in histamine shock. I F ALSINA *Compt rend soc biol* 100, 1194-1101(1929) — The changes in the alkali reserve are accompanied by a relative increase in blood Na and decrease in blood Cl. B C A

Dextrose and normohemolysis. I DUBOIS *Compt rend soc biol* 100, 1224-5(1929) — Dextrose has no influence on hemolysis by hemolysin hence the mechanism is not that in hypertonia and is not accompanied by colloidal changes of the proteins of the surface layer. B C A

Oseous exchange in fever caused by naphthylamine-yellow or by β -tetrahydro-naphthylamine. J DADLERZ AND W KOSKOWSKI *Compt rend soc biol* 100, 1241-6, 1244-40(1929) B C A

Nature of the toxic product arising from closure of the intestine. A BLANCHETIERRE AND L. SIVET *Compt rend soc biol* 101, 14-7(1929) — On occlusion of the intestine there is present a single chain substance of protease character, probably arising from the decarboxylation of leucine. It is possibly isomylamine. B C A

Action of formaldehyde, hydrogen peroxide and white phosphorus on tuberculin. J F IRVING AND C IRVING *Compt rend soc biol* 101, 153-5(1929) — Tuberculin, which is not decomposed by the above reagents, is exceptional among proteins. B C A

Affinity of hemoglobin for oxygen in anemia. G LITACZEK, H ACHERT AND J COMMERCON *Compt rend soc biol* 101, 220-1, 222-3(1929) — The affinity of hemoglobin for O is reduced in anemia. The ratio of $1/k$ to $\text{Cu} \times 10^4$, where k is the dissociation const. of the oxyhemoglobin is abnormally high in anemia. B C A

Bile protein in diabetes. J GAVRILA AND A MOGA *Compt rend soc biol* 101, 404-6(1929) — In diabetes the bile protein is increased. B C A

Susceptibility and resistance to tar cancer, an experimental study. J MATHY *Ann. Surgery* 93, 180-9(1931) — Tar cancer is the result of a local irritation as well as of a general intoxication. Certain mineral salts (Mg, Cu) used in convenient doses are capable of delaying the appearance of cancer and of lowering the percentage of induced tumors. Other salts, NaCl , salts of radioactive metals, have the opposite effect. Diet can influence the appearance and evolution of tar cancer. The liver diet activates the development of cancer. It is probable that other diets would retard its appearance. Vitamin A administered in addition to a normal alimentation has no inhibitive effect on the appearance of cancer. Preceding closely the development of tar cancer and during its evolution, the metabolism of fats is deficient and the variations of the cholesterolemia are marked. The same variations are to be noticed after repeated ionium injections. RACHLE BROWN

Dehydration in intestinal obstruction. GAVIN MILLER *Edinburgh Med J* 38, 47-53(1931) — In high simple obstruction in dogs, death is caused by loss of intestinal secretion causing dehydration and loss of Na and Cl ions. Only the blood plasma and interstitial fluid are lost. The intracellular fluid is maintained by an ionic concentration mainly of H_2PO_4 and K. These are not lost in the gastrointestinal secretion and so dehydration does not impair the intracellular fluid. RACHLE BROWN

Experiences with a high carbohydrate-low caloric diet for the treatment of diabetes mellitus. I M RADINOWITZ *Can Med Assoc J* 23, 489-94(1930) — A report from the clinical point of view is given of a diet such as the title indicates. Provided the fat content is kept very low and the cal. value is within the min. requirements, it is possible, in the majority of cases, to make the diet of the diabetic quite attractive without the use of insulin. Where insulin is necessary, the dosage and frequency of

administration are less than with diets ordinarily in use. Special articles of food are not necessary. Low plasma cholesterol and N retention in the early stages are characteristic metabolic features associated with the treatment. A. T. CAMERON

Chloropenia with excess of tissue chloride in cirrhosis. H. THIERS. *Compt rend soc Biol* 101, 475-6 (1929). *Physiol Abstracts* 15, 36, cf. C. A. 24, 5370—In alcoholic cirrhosis of the liver the following syndrome is described: (1) diminution of Cl in blood, (2) normal or raised alk. reserve, (3) urea and non-protein N diminution in blood, (4) excess of Cl in tissues. The hepatic lesion plays a part in this. It is not attributable to the same mechanism as in nephritis. H. L. D.

Superventilation tetany. FRIEDRICH HOLTZ. *Z. physiol Chem* 194, 76-80 (1931)—The tetany resulting from a prolonged period of voluntary forced breathing is accompanied by a lowering of the phosphate content of the serum. This may drop to a value as low as $1/3$ of the normal, while the Ca remains normal. A study was made of the effect on nervous excitability by measuring a low-voltage current in milliamperes required to produce tremors in the tibial nerve. No fundamental difference was observed between normal and epileptic subjects except that epileptics reached the excitable stage somewhat sooner. A. W. DOX

Disturbances in the carbohydrate metabolism in Basedow's disease. BERNHARD KUGELMANN. *Klin Wochschr* 9, 1533-4 (1930)—Thyreotoxic patients given 100 g. of levulose per os show an appreciably higher and more prolonged curve, presumably due to functional liver insufficiency. H. EAGLE

An experimental contribution to the serodiagnosis of carcinoma. I. The use of a scale-photometer in the albumin-A reaction of Kahn. W. BACHMANN. *Klin Wochschr* 9, 1540-1 (1930)—The demonstration of normal quantities of albumin-A in the serum is strong presumptive evidence for the absence of carcinoma. II. The comparison of the Kahn method of albumin-A determination by scale-photometric measurement with the modified Botelho reaction. W. BACHMANN AND FR. SCHMIDT. *Ibid* 1541-3 (1930)—A Kahn estimate of albumin A over 32 signifies the absence of cancer, readings of 32-30, with the Botelho M or Z tests negative, make carcinoma improbable, Kahn < 32, and positive Botelho M and Z tests are strong evidence for the existence of cancer, Kahn < 30, and Botelho tests negative or conflicting signify possible cancer. It is important to regulate the pH in the Botelho tests. H. EAGLE

Eclampsia and ion economy. H. ROSENVECK. *Klin Wochschr* 9, 1545-7 (1930)—In eclampsia Na leaves the blood and goes into the tissues; there is also a relative increase in muscle Ca and a decrease in K. In the liver there is an increase of Na and K and a decrease in Ca. There is a theoretical discussion as to the significance of these findings. H. EAGLE

The clinical significance of the porphyrins. HUGO KÄMMEKER. *Klin Wochschr* 9, 1656-63 (1930)—A critical analysis of the literature. H. EAGLE

Causal relationship and metabolic effects in diseases of abdominal organs and kidneys. SIGMUND FREY. *Klin Wochschr* 9, 1678-81 (1930). H. EAGLE

Organ-antibody formations in human beings. ORZDOLY FISCHER AND FELIX GEORGI. *Klin Wochschr* 9, 1817-9 (1930)—Spirochetes grown in a rabbit brain medium and injected into humans cause the appearance of antibodies which react with an alc. ext. of either spirochetes or human brain. Spirochetes grown on a rabbit kidney medium, and injected into humans, cause the appearance of antibodies against an alc. ext. of beef kidney, against sheep red blood cells, and against spirochetes. Absorption of this latter serum with sheep red cells removes its reactivity for the sheep cells as well as for the kidney, but leaves undiminished its reactivity with the spirochetes. The greater reactivity of the kidney-spirochetes antisera for spirochetes grown on a kidney medium is not affected by this absorption. F and G conclude (1) that sheep red cells, rabbit kidney and beef kidney contain common receptors, which may or may not be identical with the Forssman receptor, (2) that organ-specific antibodies may be induced in humans (anti-brain, anti-kidney), which are not species specific, (3) that the difference between spirochetes grown on kidney and brain media, manifested by their greater reactivity with the homologous antisera, are not dependent upon the organ-antibodies present in these sera. It is highly significant that none of these anti-spirochetal sera is Wassermann-positive. H. EAGLE

Decrease of the urine chloride in acute fibrous lung involvement. P. REICHEL. *Klin Wochschr* 9, 1870-2 (1930). H. EAGLE

The significance of elastin in arteriosclerosis. F. ROSENTHAL. *Klin Wochschr* 9, 1909-13 (1930)—In jaundice, there is not a uniform saturation of the body by the bile pigments, but a selective absorption determined by the affinity of the various organs for bilirubin. Thus, in dogs with experimental jaundice (chole-*medum*), despite a very high blood bilirubin,

the bilirubin content of the spleen, heart, muscle, fat, brain, etc., is minimal, while the liver, kidney, skin conjunctive heart valves, vascular intima and lungs contain large quantities. The bilirubin content of the liver and kidney is probably due to their excretory function, while the great affinity of the other tissues is due chiefly to their high content of elastic fibers. *In vitro*, these fibers absorb bilirubin from solution. In jaundiced skin, conjunctival, nuchal ligament, etc., the pigment cannot be extracted by alkali, acetone or CHCl_3 , save in small quantities, while it is readily extracted from liver or kidney. Apparently, elastin enters into a firm combination with bilirubin. It should be emphasized that the foregoing holds only for acute exsial jaundice. H. FÄLGE

The iodine-deficiency theory of goiter. C. ALEXANDER HILLWIG *Klin Wochschr.* 9, 1914-6 (1930).—See C. A. 24, 3052. H. FÄLGE

Icterus neonatorum gravis. H. KLEINSCHMIDT *Klin Wochschr.* 9, 1951-4 (1930).—The disease is characterized by a high degree of erythroblastosis. The large quantity of hematin in the blood and the pronounced siderosis both indicate toxic destruction of the blood. Early transfusion is suggested. H. FÄLGE

Acid-base balance in gastric and duodenal ulcer. F. REICHLE AND F. FREYBURGER *Klin Wochschr.* 9, 1954-9 (1930).—The authors cannot confirm the findings of Bahnt that there is acidosis in cases of duodenal ulcer. H. FÄLGE

Testing the detoxicating function of the liver. NICH. SCHOUR AND CEL. ROSEN-GARTEN *Klin Wochschr.* 9, 1968 (1930).—With increased putrefaction and a normally functioning liver, there may be indicanuria, but indole or uroscopine does not appear. In cases of liver insufficiency, indican decreases, and the other 2 substances appear in the urine, the quantity depending upon the degree of intestinal putrefaction and of liver damage. H. FÄLGE

Studies in the experimental autohemolysis of rabbits. SHUNZO SUMAMI *Tôhoku J. Exptl Med.* 16, 277-97 (1930).—A mixt. of lecithin and pig serum injected intravenously into rabbits causes the production of autohemolysis. A certain proportion of rabbits normally contain autohemolysis in the serum, and only a certain proportion contain the necessary receptor groups in the red cells which can combine with the antibody. H. FÄLGE

The distribution of water in the animal body. I. Distribution of water in the rabbit with a functional impairment of the thyroid gland. HIYOS. KUSAKIRI AND KENJIRO TAKEDA *Tôhoku J. Exptl Med.* 16, 329-40 (1930).—In hyperthyroidism, the body as a whole loses water, particularly the skin, in hypothyroidism there is water retention, also chiefly in the skin. Variations in the other organs are not constant. H. FÄLGE

Studies in the colloid-osmotic pressure of the blood in normal and pathological conditions. III. Colloid-osmotic pressure of the blood in hypotension and hypertension. FUSAKICHI NAKAZAWA AND JIRO IZUMI *Tôhoku J. Exptl Med.* 16, 341-51 (1930), cf. C. A. 24, 2801.—In both hypotension and hypertension without kidney involvement there is no significant change in the colloid-osmotic pressure of the blood or in the osmotic pressure per gram of serum protein. In hypertension with kidney damage, however, the total colloid osmotic pressure is very low, even though the serum protein may show no striking change. IV. Colloid-osmotic pressure of the blood in experimental kidney damage. KOTARO KIMURA AND FUSAKICHI NAKAZAWA *Ibid.* 352-60.—Following poisoning with cantharidin or U, the total colloid-osmotic pressure of dog serum is strikingly low, as is also the pressure per unit serum. In marked contrast following U poisoning, the colloid osmotic pressure of the urine per unit protein is 3-5 times its normal value. Nephrectomy or bilateral obstruction of the ureters has qualitatively the same effect upon serum colloid pressure as poisoning, although to a less degree. These results are ascribed to the presence in serum of larger protein aggregates, due either (1) to an increased capillary permeability allowing the passage of small protein "molecules" into the tissues, (2) albuminuria, with passage of the smaller protein particles into the urine; (3) functional changes in the organs responsible for the regulation of blood protein. These results take on an added significance in the light of the recent work of Van Slyke. H. FÄLGE

Metabolism of blood glucose in experimental trypanosomiasis. G. ZOTTA AND E. RADACOVICI *Arch. roumaines path. exptl. microbiol.* 2, 55 (1929), *Rev. hyg. méd. prév.* 53, 65 (1931).—Exptl. studies with guinea pigs in Nagana showed no definite relation between hyperglucemia and the nos. of trypanosomes present in the blood. It is believed that the hyperglucemia in trypanosomiasis is due to a disturbance of the central gluco-regulating mechanism in the body. C. R. FELLERS

Cholesterol content, viscosity and pH of blood after splenectomy. L. BUGNARD *Compt. rend. soc. biol.* 101, 546-8 (1929); *Physiol. Abstracts* 14, 661; cf. C. A. 24, 3556.—

Zn was demonstrated as the cause in certain cases. In the intervals between crises the Zn in the urine amounted to 1 mg. per 24 hrs., whereas during the attacks it was at least 3 times more abundant. The crisis is regarded as probably due to zincopeny and an indication is presented for treatment with Zn. G. G.

Mode of action of manganese salts in immunization. A. L. MUIR AND R. H. GRANT. *Compt. rend. ex. biol.* 103, 231-3 (1930), *Physiol. Abstracts* 15, 132.—Injection of Mn salts in the dog caused a marked leucocytosis. It is probable that the increase of immunity produced by this method is due to this leucocytosis. Substances capable of stimulating the leucogenic function should be made use of for immunizing purposes. G. G.

Variations of affinity of hemoglobin for oxygen in hyperthyroidism. G. LITVACK, H. ACHUT AND J. COOMERICO. *Compt. rend. ex. biol.* 102, 1310-1 (1930), *Physiol. Abstracts* 15, 89.—In hyperthyroidism an increase of the const. k , expressing the affinity of hemoglobin for O was found. An increase of k corresponds to a lower oxygen curve and to a more ample liberation of O to the tissues. The increase is not so accentuated as in anemic subjects. G. G.

The pathology of iron. I. PRINGSHEIM. *J. Lab. clin. Med.* 13, 807-18 (1928), *Physiol. Abstracts* 13, 541.—A histological study of blood of brain and of other tissues in cases of pernicious anemia and malignant growths was carried out with the Berlin blue and methyl green methods of staining. Many speculative suggestions are put forward to account for the distribution of Fe found by this method. G. G.

Ammoniacal output in alkalosis. M. BUCOSOVSKI, P. BOUTANIER AND H. WARRINER. *Compt. rend. ex. biol.* 100, 207-9 (1929), *Physiol. Abstracts* 14, 39, cf. C. A. 23, 4970.—It has not been possible to confirm the views of James that increased NH_3 in the urine is anti acid in function and that excess of NaHCO_3 by the mouth annuls the NH_3 eliminated. NH_3 seems to be taken from the blood by the kidney in obedience to a law similar to that governing the secretion of urea. It is further confirmed that the NH_3 output does not follow that of total N, and exptl. alkalosis causes reduction of the ammoniacal excret. which, according to the law of Hasselbalch, should remain const. G. G.

Blood creatinine and creatine in pathological states. J. GAVRILA, V. VIMR AND R. RAMRANTY. *Compt. rend. ex. biol.* 100, 381-3 (1929), *Physiol. Abstracts* 14, 471.—In 15 cardiovascular cases the creatinine varied from 1.3 to 2.14 and creatine from 3.04 to 0.65 mg. %. In 13 cases of diabetes the creatinine varied from 1.33 to 2.5 and creatine from 4.05 to 7.5 mg. %. In leucemia creatinine varied from 1.42 to 2.5 and creatine 3.6 to 0.72 mg. %. In 21 cases of chronic nephritis the creatinine varied from 1.33 to 10 and creatine from 3.93 to 17 mg. %. N showed parallel variations. In 7 cases of pulmonary tuberculosis the creatinine varied from 1.26 to 1.66 and creatine from 3.44 to 5.81 mg. %. In tuberculosis, pleurisy and bronchiectasis the normal limits were maintained. G. G.

Soap-venom complexes in immunization against cobra venom. M. REVEN. *Compt. rend. ex. biol.* 103, 143-4 (1930), *Physiol. Abstracts* 15, 197.—Venom of the cobra mixed with Na oleate or palmitate loses its toxicity. The immunising power of this mixt. was tested by repeated injection into guinea pigs. In 80% of animals it was distinctly demonstrated. Soap behaves toward venom as it does toward toxin, notwithstanding the remarkable differences of chem. structure between venom and toxin. It annuls the toxic action while leaving the antigenic properties intact. E. C. K.

Pathological variations of protein osmotic pressure and the protein composition of blood serum. CH. ACHARD, A. GAGNIER AND J. COUDANE. *Bull. soc. chim. biol.* 12, 417-40 (1930)—By using the osmometer of Grigaut and Bouteux, greater variations were found in the osmotic pressure of the serum proteins than in the ratio of albumin to globulin. From the study of 4 normal and 50 pathol. subjects very little correlation was found between the values and the evolution of disease, except the low total protein in nephritis. The albumin globulin ratio was found to be of little significance. G. G.

The iodine content of the thyroid gland accompanying simple and exophthalmic goiter. P. DANTON AND H. SCHWENK. *Bull. soc. chim. biol.* 12, 773-6 (1930)—The av. I contents from 13 cases of simple goiter and 11 cases of exophthalmic goiter were 0.003 (0.010-0.215) and 0.003 (0.010-0.267)% dry wt., resp. G. G.

The absence of tryptophan and histidine as the cause of pernicious anemia. Therapy of the anemias by supplementing the organism with hemalogenic amino acids. GEORGE POSTS AND LUCY THURMAN. *Compt. rend. ex. biol.* 101, 1203-7 (1930)—P. and T. argue that the absence of tryptophan and histidine may be regarded as the cause of anemias. Experimentally the administration of the 2 amino acids in cases of leu-

cemic, secondary, splenomegalic and neoplastic anemias produced encouraging results

B. S. LEVINE

The "zoning" phenomenon in complement fixation with cholesterolized alcoholic beef heart extract. Its mechanism and significance. B. S. LEVINE. *J. Infectious Diseases* 48, 189-202(1931)—L. presents theoretical considerations and exptl. evidence to explain the mechanism and significance of the "zoning" phenomenon occurring in complement fixation employed in the lab. diagnosis of syphilis in which the preliminary incubation is done in the ice box. By "zoning" L. means the occurrence of a stronger fixation of complement in a lesser amt. of suspected serum. He concludes that in the serum of most adults there are various types of immune substances either in the free state or in union with their respective homologous antigens. Upon the addn. of cholesterolized alc. beef heart suspension, the action of which L. proves to be non-specific, interfaces are created at the suspended particles of the colloid. The so-called immune complex concentrates at such interfaces in spherical and spheroid configurations. With the increase in the diln. of the serum, the dispersion of the suspended complexes increases, leading to an increase in the surface area of the sensitized spheres in accordance with the following formula: $S_2/S_1 = \sqrt{n_1/a_1}$, where S_2 = surface area of the suspensoid in the more dil. serum, S_1 = surface areas of the sensitized spheres in the original diln. of the serum, n_1 = the no. of suspended spheres in the more dil. serum, and n_1 = the no. of spheres in the serum of the original diln. In a special case where $n_1 = 1$, the formula becomes $S_2/S_1 = \sqrt{a_1}$. With a ration of the dispersion in the serum of the higher diln. to that of the lower exceeding 8:1 the consumption of the complement by vol. $v/2$, of full strength serum will be greater than the consumption of complement by vol. v , of full strength serum when both are dild. with the saline antigen suspension to the same final volume. Hence, "zoning" will appear in all serums in which the consumption of complement by the so called "non specific" antibodies in the first tube of the cold fixation procedure just exceeds, equals or nearly equals one "unit." The significance of the "zoning" phenomenon rests on the fact that it definitely proves the general immunologic rather than the specific leucic nature of complement fixation with cholesterolized alc. beef heart ext. L. concludes that to consider serum specimens manifesting "zoning" "positive for syphilis" is paramount to denoting positive for syphilis specimens showing hemolytic inhibition in a complement fixation procedure where 10 or 11 unit of complement is considered the "dose." B. S. LEVINE

Chemical investigations in Dercum's disease. JAVINE H. PAGE. *Virchow's Arch. Path. Anat.* 279, 262-4(1930)—The depot fat of the abdominal cavity and that of the thigh were similar, and likewise normal, in a case of Dercum's disease. Chem. findings do not explain why fat is laid down in Dercum's disease. E. R. LONG

Animal ochronosis and porphyry. R. TRENTSCHER. *Virchow's Arch. Path. Anat.* 279, 731-8(1931)—Beef kidneys, bone marrow and bone from a case of so-called "animal ochronosis" were examd. for the presence of porphyry. The results were negative in the kidneys and bone marrow and positive in the bone. In bone the substance was detected by direct chem. analysis and spectroscopic examn. of extd. pigment. F. concludes that the ochronosis, or better, osteohemochromatosis of cattle depends, in part at least, on porphyry. E. R. LONG

The glycogen content of the cadaver liver. HANS POPPER AND OSEAR WOJASEK. *Virchow's Arch. Path. Anat.* 279, 819-68(1931)—The total carbohydrate and glycogen content of the liver was detd. in 177 human bodies. Examn. was made as soon as possible after death. Usually the values obtained by chem. detn. of total carbohydrate and histol. estn. of glycogen corresponded. In sudden deaths, in persons previously in good health, the total carbohydrate content of the liver ranged from 1.56 to 6.17% of the moist wt., with considerable amounts of glycogen histologically visible. When death occurred after a wasting disease, the figure ranged from 0.24 to 1.53%, with little or no glycogen demonstrable in sections. In 22 cases of cirrhosis of the liver, where death came gradually, the value for total carbohydrate was 0.38-0.94%. Histologically demonstrable glycogen was relatively high. In 17 diabetes the total carbohydrate was 1.19-8.50%, $s. e.$, significantly higher than in other pathological conditions and equal to, or higher than, in the normal state. Glycogen could always be found, usually in considerable amount. In 11 liver tumor metastases total carbohydrate was 0.28-1.45%, glycogen was histologically not demonstrable. In 26 kidneys the total carbohydrate ranged from 0.28 to 1.64%, the higher values being found in cases of diabetes mellitus. E. R. LONG

A study of the blood urea clearances with relation to diuresis in normal and nephritic animals. ROBERT L. JOHNSTON. *J. Lab. Clin. Med.* 15, 943-52(1930)—The ability

of the kidneys of the normal and nephritic, fasting animal to clear the blood of urea is correlated with the water-diuresis curve. A simple diuresis test is described for the determination of the status of renal function. F. R. MAIN.

The incidence of lipids in urine. A report of the micropolariscopic examination of 1470 specimens. ARTHUR T. BRICE, JR. *J. Lab. Clin. Med.* 15, 933-40 (1930).—Lipoid substances were detected in 13.5% of the pathol. urines examd. A high incidence of urine lipoids assocd. with degenerative diseases of the kidneys suggests that pathol. conditions in the gall bladder may be an etiologic factor. F. R. MAIN.

The Congo red test with special reference to excretion of the dye in the urine. NELSON W. BAKER AND ALBERT M. SNELL. *J. Lab. Clin. Med.* 16, 202-70 (1931).—The disappearance of more than 60% of intravenously injected Congo red from the blood stream in 1 hr. appears to be a sp. test for amyloidosis. The dye is not excreted in the urine in the absence of albuminuria. It is excreted in relatively large amounts in chronic lipid nephrosis and in some instances, in nephrosclerosis with albuminuria and in renal amyloidosis. F. R. MAIN.

Adrenaline glucemia in disturbances of the liver. G. NANNINI. *Minerva med.* 1930, II, 574-7.—The glucemic curve was studied after the injection of adrenaline into 20 patients suffering from various hepatic lesions. The results indicate that the glucemic values are useful in the diagnosis of hepatic lesions and also for differentiating various types of icterus. PETER MASCOCC.

The function of the liver in cardiopathic cases investigated by means of the amino acid curve. LUCIO ALZOLA. *Minerva med.* 1930, II, 917-23.—The hepatic function in patients with various forms of cardiac disturbances was studied by detg. the amt. of amino acids in the blood after the intravenous injection of glycocoll. A. concludes that the deaminizing action of the liver in cardiopathic cases varies according to the valvular lesion, the relation between compensation and the deaminizing power of the liver is not const., in aortic lues, glycocoll disappears from the blood with more difficulty; as a general rule, the older the lesion, the more impaired is the hepatic function. PETER MASCOCC.

The determining factors of the hypoglucemia in diabetic children. HYDRO B. LANDAUER. *Semana med.* (Buenos Aires) 1931, I, 100-4.—In 50% of diabetic children subjected to a diet and insulin treatment, hypoglucemic symptoms develop. This is due to a high variability of the glucemia in children. Even doses of 2 to 5 units may cause the syndrome which is not controlled by the sugar forming substances in the diet. The symptoms appear at a blood sugar between 0.07 and 0.024%. A direct relation between the sugar level and the severity of the symptoms could not be established. A. E. MEYER.

The passage of some normal antibodies from the mother to the fetus. ELISA MORELLI. *Sperimentale* 34, 431-41 (1930).—The bactericidal power against cholera vibrios in the blood of the pregnant guinea pig is not transmitted to the fetus. The hemolytic capacity of rabbit blood decreases during pregnancy and is not present in the new born. The lack of antibodies does not seem to be connected with changes of the complement in the blood. A. E. MEYER.

The action of proteins in the growth of grafts of homologous neoplasms. GIOVANNI FAVILLI. *Sperimentale* 34, 489-504 (1930).—Mouse carcinoma and rat sarcoma were transplanted after a previous stimulation of the reticulo-histocyte system by injection of protein. The growth was stimulated in this case, this is probably due to injuries in the spleen caused by the proteins. A. E. MEYER.

Isolation of the toxin of Fraenkel's gas gangrene bacillus. II. ALFRED EDEL. *Buchern. Z.* 229, 265-8 (1933); cf. C. A. 24, 5854.—Culture filtrate acid to litmus were evapd. *in vacuo* and completely pptd. by the addn. of 96% alc. The alc. soln. was then evapd. *in vacuo* to a very small vol. and again pptd. with alc. The alc. soln. of the poison was treated once more as above, until the condensed material dissolves completely in 96% alc. Upon the addn. of abs. alc., however, a thick, inactive ppt. is again produced. The soln. of the poison in the abs. alc. is now put through the same process of purification with abs. alc. until the concentrate completely dissolves in the abs. alc. and gives no ppt. even with an alc. soln. of Pb(OAc)₂. The soln. of the poison in abs. alc. is subjected to electro-dialysis against distd. water, when only the cathode portion reveals strong activity and an alk. reaction. This portion was greatly concd. *in vacuo* and mixed with 4 vols. of 96% alc.; upon standing in the ice box inactive crystals sep. out. The alc. soln. is again concd. and purified by electro-dialysis. The alk. dialyzate is pptd. with Pb acetate, the soln. freed from Pb and once more electro-dialyzed. The dialyzate is concd. *in vacuo* and treated with alc. until the residue of the abs. alc. soln. is completely sol. in abs. alc. Pptd. now with ether, the

ether alc soln retains the bulk of the poison. A further attempt is in progress to isolate the poison from this ether alc soln. S. MORCULIS

Studies on glucolysis of the blood. L. H. K. BARRENSCHIEEN AND KARL HÖBNER. *Biochem Z* 229, 329-42(1930).—An investigation of a number of glucolysis curves shows that the application of the equation of a first order reaction even with Fukushima's corrections for adsorption and diminished enzyme activity does not yield uniform reaction consts, more particularly during the initial stages of glucolysis. Studies on blood from normal fasting persons show that there is a 15-30 min induction period which can be eliminated under the influence of SO_4 ions. The blood glucolysis is similarly stimulated by PO_4 and AsO_4 ions. The inorg P during the first hrs of glucolysis does generally diminish; this is attributed to a process of phosphorylation, the SO_4 ions stimulating this diminution in the P. It is interpreted that SO_4 ions activate the P esterifying process. Previous excessive ingestion of carbohydrate causes the disappearance of the induction time observed in the glucolysis of blood from the fasting subject, and the decrease in inorg P occurs sooner and to a greater extent. Similarly the addn of hexosediphosphate (but not of mono-phosphate esters) causes the disappearance of the induction time and accelerates the glucolysis. S. MORCULIS

Disturbances in the mineral regulation mechanism in diseases of cattle. A study of tetany. B. SJÖLLFMA AND L. SEEKLES. *Biochem Z* 229, 358-80(1930).—A study of blood serum of cattle suffering from delivery paresis and a "grass tetany" reveals some remarkable changes in the mineral metabolism, particularly in the Mg content, the concn of Ca ions and of the inorg P. In the former disease these quantities were on the av. 2.19, 0.44 and 2.16 mg %, in the latter, 0.46, 1.18 and 4.33 mg %, resp. The difference between these 2 pathol states was also manifested in the total Ca content, namely, 4.35 and 6.65 mg %, resp. In the normal cattle the concns of Mg, P, Ca^{++} and Ca on the av. were 1.66, 4.57, 1.65 and 9.35 mg %, resp. A graphic method has been developed for representing the total blood content of Mg, Ca and P in such a way that the normal and pathol states are spatially sep'd. In delivery paresis the ratio $\text{Ca/Mg} = 2$ approaches that (0.9) existing when narcosis is produced through the injection of a Mg salt. In "grass tetany" this ratio is $\text{Ca/Mg} = 14.6$, while in normal cattle it is 5.0, showing that there is actually a condition of excitation. In "grass tetany" the Mg content of the blood varies according to the time of the collection of the sample. If taken at the beginning of an attack of tetany the Mg is low and after the attack it is almost invariably high, possibly because of the strong muscle contraction during the convulsions. The tetany attacks in "grass tetany" occur even when the Ca/P is relatively high, namely 1.5, whereas in parathyroid tetany the attack does not take place unless this ratio is less than 1. In both pathol conditions however, the % of diffusible Ca undergoes wide variations. Whereas in normal cattle the diffusible Ca constituted 48-66% of the total, it ranged from 25 to 100% during delivery paresis and from 40 to 100% during "grass tetany," but no connection was observed between the change in the % of diffusible Ca and the attacks. The inorg P of the serum is completely diffusible under all conditions. The diffusible fraction of the Mg varied greatly in the normal and in the sick cattle. In delivery paresis no evidence was found of a relationship between the concns of the Ca^{++} , H^+ , HCO_3^- and HPO_4^{--} of the serum. S. MORCULIS

Behavior of blood water during asphyxia. TAKESHI SASO. *J Biochem (Japan)* 12, 281-315(1930).—The plasma vol in acute asphyxia decreases about 14.8%, this being practically independent of the dehydration of the blood. The sp gr of the serum is always raised by about 0.3% in acute asphyxia. The dehydration of both blood and serum during acute asphyxia is very small, being on the av. 0.6% for serum and only 0.3% for the whole blood. In prolonged asphyxia, on the contrary, the water content of the serum or blood always increases and there is never any dehydration. The dehydration occurring in acute asphyxia is not extensive enough to account for all the changes in the concn of the various blood constituents, and the passage into the blood from the tissues must be the chief source from which these increases originate. Particularly in prolonged asphyxia, when there is actual hydrema, the increased concn of different blood constituents is obviously due to migration of substances from the tissues. S. MORCULIS

Comparative study of proteins in blood serum and pathological serosities. CH. ACHARD AND A. ARCAN. *Compt rend* 189, 510-2(1929).—The authors question the fact that the proteins of exudates and transudates come only from blood plasma. Pointing out that the protein content of plasma varies notably in pathol. conditions they compare the protein content of blood serum with that of serous liquids. In 12 cases of pleurisy (10 inflammatory and 2 cardiac) about 50% of protein was found in

the serous liquids while 63-86% was found in the blood serum. The ratios between the proteins of blood plasma and serous liquids (serum/liquid) were 1.25 to 1.76 for the inflammatory cases, and 2 to 2.89 for the cardiac. The ratios for individual proteins were albumin 1.16 to 2.24 (one exceptional value of 0.91), globulin, 1.14 to 2.23 (2 exceptional values of 0.96 and 4.13), *myxoprotein*, 1.07 to 1.79 in six cases, but 3.53, 5.66, 6.50 and 17.92 in four cases. In the two cardiac cases the values were raised for albumin and globulin, but were discordant for *myxoprotein* (1.07 to 5.66), in two cases exceeding the ratio for globulin in the blood serum but not in the serous liquid. In seven cases of dropsy the values of serum/liquid were 1.13 to 2.96 (one exceptional value of 1.60 in a case of serum rich in protein). The ratios for the individual proteins were albumin 3.61 to 46.47, globulin, 1.60 to 40.13, *myxoprotein*, 4.75 to 32.31. The ratios show that there was always much less albumin, globulin and *myxoprotein* in the serous liquids than in the blood serum, but the proportion was much more variable for *myxoprotein*. A and A conclude that most of the proteins of serous liquids come from blood plasma, but that this does not exclude the possibility of protein formation in the serous membranes. A. L. RAWLIN.

Guanidine-like substances in the blood in experimental yellow fever. G. M. FINDLAY AND L. HINDLE. *Lancet* 1930, II, 678-9.—In exptl yellow fever in monkeys there is an increase in the blood content of guanidine like substances, similar to that associated with other conditions involving liver necrosis. This increase can be reduced by the administration of Ca lactate, which, however, does not prevent death but does tend to prevent hemorrhages. The method of Pflüger and Myers was used for detg. the guanidine like substances. F. B. SINFERT.

Manganese in foodstuffs and its possible relation to cirrhosis of the liver. A. I. BOYCOFF AND G. R. CAMFORD. *Lancet* 1930, II, 959.—The method of Bradley (Lindow and Peterson C. A. 22, 125) was used for detg. the Mn. Various food products were analyzed and of them cloves, tea, pepper, ginger, coriander, winkles and scallops contained the most Mn. F. B. SINFERT.

Plasma-protein loss with edema but without proteinuria and its bearing on the concept of nephrosis. T. I. BENNETT, E. C. DODDS AND J. D. ROBERTSON. *Lancet* 1930, II, 1006-8.—Massive loss of protein may occur from the blood by routes other than the kidney and edema follows this loss. In such cases no pathol. changes are produced in the kidneys. F. B. SINFERT.

Urea content of the cerebrospinal fluid in status epilepticus. O. R. BIRSON AND J. TAYLOR FOX. *Lancet* 1930, II, 1233-4.—There was a high urea and glucose content of the cerebrospinal fluid. F. B. SINFERT.

Inflammation. V. The mechanism of fixation by the inflammatory reaction. VALY MENKIN. *J. Exptl. Med.* 53, 171-7(1931); cf. C. A. 24, 4823.—Microscopic studies show the presence of a network of fibrin within the tissues and numerous thrombosed lymphatics at the site of inflammation. Pyrid. Fe compds., possibly coagulated horse serum or particulate matter caught in this fibrinous reticulum, will disseminate less readily than trypan blue from the site of inflammation. Trypan blue injected at the periphery of an inflamed area fails to enter the site of inflammation. Fixation of foreign substances by the inflammatory reaction is primarily due to mechanical obstruction caused by a network of fibrin and by thrombosed lymphatics at the site of the inflammation. VI. Fixation of trypan blue in inflamed areas of frogs. *Ibid* 179-83.—Trypan blue, injected into the circulating blood stream of frogs, accumulates rapidly in an inflamed area; injected directly into the area of inflammation, it is fixed *in situ* and fails to diffuse outward. If the dye is injected at the periphery of an inflamed area, it fails to enter the site of the inflammation. This failure of penetration is caused by the occlusion of lymphatic vessels and the presence of a network of fibrin in the inflamed area. C. J. WEST.

Mineral content of pastures research (ASTON) 12. The detection of horse meat as an adulterant in sausage and other studies of the precipitin test (BOLIN) 12.

HOLMES, BARBARA: Cancer and Scientific Research. New York: The Macmillan Co. 160 pp.

H—PHARMACOLOGY

A. N. RICHARDS

Elimination of quinine in the bile. P. CAUJOLLE. *Bull. soc. chim. biol.* 12, 299-300 (1930).—The presence of quinine in the bile of dogs, after intravenous injection of an isotonic soln. of the neutral sulfate, is demonstrated by observing the fluorescence in

H_2SO_4 when exposed to light from a Hg vapor lamp after passage through a suitable filter. Elimination is more rapid in the urine than in the bile. B C A

Intestinal elimination of iron in the dog. V. HEVAIGUËS AND A. ROCHE. *Bull soc chim. biol* 12, 404-8(1930).—A daily oral administration of Fe (300-1675 mg) as lactate does not affect elimination by the colon. The intestinal absorption of Fe (detd. by the $TiCl_3$ method) varies with the salt used, and the results do not support the theory of absorption in the small intestine and elimination by the large intestine. B C A

Arrest of the internal secretion of the pancreas during decamethylenediguanidine hypoglycemia. E. ZUNZ AND J. LA BASAR. *Compt rend soc biol* 101, 141-3(1929), cf. C. A. 24, 1149.—Synthalin hypoglycemia is accompanied by a diminution of the secretion of insulin. B C A

Chemistry of overventilation. G. POPOVICU AND H. POPESCU. *Compt rend soc biol* 101, 406-8(1929).—In convulsions caused by over ventilation there is an increase in serum Ca and a decrease in the serum phosphate. The decrease in phosphate is reduced by ergotamine and increased by adrenaline and ephedrine. All these drugs tend to reduce the serum Ca. B C A

Effect of yohimbine on blood sugar. S. HANSON. *Compt rend soc biol* 101, 601-3, 603-5, 605-7(1929).—Yohimbine causes a slight rise in blood sugar, in small doses it suppresses adrenaline hyperglycemia although with larger doses a rise is observed. The increase by yohimbine of the hypoglycemic action of insulin is discussed. B C A

Effect of adrenaline on the respiratory metabolism in exophthalmic goiter. I. SCHILL. *Magyar Orvosi Arch* 30, 202-0(1929). B C A

Influence of electrolytes on the syneresis and clotting of blood. SATYA PRAKASH AND N. R. DHAR. *J Phys Chem* 35, 629-37(1931).—The electrolytes used were KCl , $CaCl_2$, $(NH_4)_2SO_4$, KF , $K_2C_2O_4$, $AcONa$, $Na_2C_2H_3O_2$, $NaOH$ and Na citrate. Each test sample contained 230 cc of goat blood and was added to 250 cc by the addition of H_2O and salt soln. The smallest quantity of electrolyte used was 2 cc of N soln and the largest was 15 cc. of 3 N soln. In all cases syneresis decreases with increasing concn of any one electrolyte. The concns necessary to suppress syneresis in the clots formed were: KCl 0.12 N , $CaCl_2$ 0.0845 N , $(NH_4)_2SO_4$ 0.176 N , $K_2C_2O_4$ 0.016 N , Na citrate 0.016 N , KF 0.049 N , $NaOH$ 0.039 N . If still higher concns. of electrolytes were added the blood was stabilized and no clotting occurred. Stabilizing of blood by fluorides, citrates and oxalates is due not to removal of Ca but to the influence of the anions. Clotting of blood and syneresis are affected by the same forces which affect formation and syneresis of both org. and inorg. gels. F E BROWN

Therapy of spontaneous mouse cancer. Failure of tuberculin, karkinolysin and some inorganic compounds therein. BURTON T. SIMPSON AND MILLARD C. MARSH. *Ann Surgery* 93, 169-79(1931).—Fifty-six inorg. compds. contg. 33 chem. elements for which there is evidence indicating therapeutic value for transplantable cancer and 4 forms of tuberculin, and 1 ext. of thymus have been found to have no therapeutic action on spontaneous mouse cancer. RACHEL BROWN

Experimental and clinical studies on the treatment of cancer by dichloroethyl sulfide (mustard gas). FRANK E. ADAIR AND HALSEY J. BAGO. *Ann Surgery* 93, 190-9(1931).—Mustard gas soln. is an effective agent against a localized cancer lesion. RACHEL BROWN

Excretion urography. An experimental investigation of the properties of uroselectan. W. ARTHUR MACEY. *Glasgow Med J* 34, 9-18(1931).—Uroselectan, the Na salt of 5-iodo-2-keto-1-pyridineacetic acid, is well tolerated by rabbits even when administered in amts. proportionately 10 times the human dose, provided the injection is made slowly. Administered rapidly, doses smaller than the max. may be fatal. There is no evidence that the kidneys are injured during the excretion of the concd. soln. of uroselectan. The urinary tract of the rabbit may be demonstrated by x-rays during the excretion period. R B

The results of the use of acriflavine hydrochloride in the treatment of undulant fever. D. S. TWONZA. *Can Med Assoc J* 25, 685-8(1931).—Results from a number of cases suggest the intravenous use of acriflavine materially shortens the course of undulant fever, and prevents the development of an incapacitating arthritis. A T CAMERON

The danger of indiscriminate iodine prophylaxis. J. F. REITH. *Pharm Weekblad* 68, 77-82(1931).—A review of the arguments against the indiscriminate and unsupervised use of iodized salt. A W. DOX

The silica content of human blood and its alteration by administration of silica

HEINRICH KRAUT *Z. physiol. Chem.* 194, 81 (1931). The SiO_2 content of the blood of normal persons varies between 1 and 3% of the solid matter, but each individual maintains his SiO_2 value within much narrower limits. The blood of tuberculous subjects contains somewhat more SiO_2 than that of normal persons. Administration of easily absorbed SiO_2 increases the SiO_2 content of the blood by a multiple of the quantity administered. After discontinuing the treatment several weeks elapse before the SiO_2 content sinks again to normal. Two methods of administration were employed: (1) oral administration of silicic acid, a glycol ester of silicic acid, and (2) inhalation of a spray containing a low mol. silica not obtained by hydrolysis of SiCl_4 or by treatment of silicates with acids at pH 2-4. The 2nd method is much more efficient. The av. of 27 cases showed an increase of 1.6% over the normal, while in 1 case it rose to 4.5%. The inhalation experiments show that the increase does not correspond to the quantity administered, but that a mobilization of the SiO_2 of the organism occurs, so that the SiO_2 value of the blood remains altered for a considerable time. Eventually the original value returns. The far greater absorption by inhalation is explained by the fact that the orally administered silicic acid encounters an unfavorable pH in the intestine and the bulk of it is there converted into an acid and difficultly absorbed form. A. W. D.

The fate of ternary chains other than those with three atoms of carbon. V. ZASANI *Ann. physiol. pharmacol.* 4, 163 (1930). The subcutaneous administration of AcOH to rabbits causes an increased heat production equal to the energy produced by its complete oxidation. Entrance of ethylene glycol has no such effect. *Z.* concludes that the metabolism of proteins or of carbohydrates does not necessarily involve the formation of ternary C chains as an intermediate product of metabolism. H. I.

The significance of chemical configuration for the pharmacological effects of adrenaline-like substances. HISAKI HASAMA *Arch. exp. Path. Pharmacol.* 153, 161 (1930) - β -phenethylamine, the basic substance of the adrenaline group, has only slight sympathomimetic action. The introduction of a phenolic OH group in the para position or of an alcoholic OH group in the β C of the side chain increases its activity. Interchanging the NH and the OH groups of the β hydroxy β -phenethylamine destroys its sympathomimetic action. Increasing the length of the side chain by a Me group decreases the activity by $1/2$. The lability of the adrenaline groups is not necessarily held by the OH groups of the benzene ring. Replacing a H in the α C by a second phenyl group destroys the activity upon nerves, but endows the substance with strong muscle stimulating properties. Oxidizing the secondary amine to a ketone diminishes the activity. The toxicity of the various members of the group does not parallel the intensity of the sympathomimetic action. H. FAGEL.

The kidney action of piloselin. Investigations upon frog kidneys. HANS HAUSLER *Arch. exp. Path. Pharmacol.* 153, 167-213 (1930) - With a glucose free perfusing fluid (Mines soln.) piloselin does not cause any change in urine secretion. Its glomerular action is due (1) to an inhibition of the normal tubular reabsorption, and (2) to a change in the permeability of the vessel walls for glucose, which allows its passage through the walls of the tubules into the urine. H. FAGEL.

Distribution of insulin in the organism following its injection. L. HANS HONSTERS *Arch. exp. Path. Pharmacol.* 153, 211-30 (1930). Following its intravenous injection into rabbits insulin is stored primarily in the liver. Within 60 min., however, it appears in the muscles, and gradually increases in quantity 50% of the total injected being recovered after 3 1/2 hrs. from the muscles. In marked contrast, following subcutaneous injection, the max. storage is reached in 60 min., and then only amounts to 30% of that injected. *H.* raises the question as to whether the injection of insulin stimulates an endogenous secretion. II. Relationship between method of nutrition and blood sugar. *Ibid.* 211-11 - In both alimentary and adrenalectomy hyperglycemia the quantity of insulin which can be extracted from the organs is increased. Animals on a carbohydrate rich diet also show increased insulin. In such animals, the administration of insulin causes a mobilization of endogenous insulin in the tissues which may exceed the quantity injected. Exogenous insulin is much easier to extract than endogenous. *H.* suggests that there are 2 types of insulin, fixed tissue insulin and free, loosely bound pancreatic insulin. In the diabetic the effective tissue insulin would be decreased. H. I.

Pharmacological and physiological studies on the sweat centers. IV. The effect of inorganic cations upon the thermal and sweat center in the midbrain. HISAKI HASAMA, *Arch. exp. Path. Pharmacol.* 153, 291-308 (1930), cf. C. A. 24, 2184-K, 1a and Na ions injected subcutaneously into the tuber cinereum cause an increase in temp. and in sweat secretion, Mg and Ca ions are inhibitory. H. FAGEL.

The mechanism of the lowering of blood pressure by histamine. MAX HUCHTENBACH AND RUDOLF MUELLER *Arch. exp. Path. Pharmacol.* 153, 309-21 (1930) - In cattle dogs

and rabbits very small quantities of histamine cause capillary dilatation acceleration of circulation, moderate increase of arterial pressure and occasionally also of venous pressure. All these effects can be explained by either a decrease in the total vessel vol. or an increased blood vol. Larger doses cause a transient hypertension, followed by hypotension. Both arteries and veins contract, particularly the latter, causing a diminished flow of blood through the dilated capillaries into the veins. The fall in arterial pressure is therefore due to imperfect filling. There is occasionally an increased liver vol., possibly because of the filling of the hepatic capillaries with blood. A temporary decrease in respiration is followed by an acceleration, possibly on account of the increased pressure in the right auricle. II EAGLE

The effect of homotropine methonitrate upon gastric secretion. M. TRUNFENBAUM *Arch. exp. Path. Pharmacol.* 153, 325-30 (1930).—Homotropine methonitrate is only $1/10$ as inhibitory upon gastric secretion by dogs as is atropine. II EAGLE

The effect of an extract of the posterior lobe of the hypophysis, morphine and caffeine upon the activity of the kidneys. B. SAGER. *Arch. exp. Path. Pharmacol.* 153, 331-40 (1930). II EAGLE

The reaction of snake hearts to certain cardiac poisons. OTTO GESSNER. *Arch. exp. Path. Pharmacol.* 153, 347-58 (1930), cf. *C. A.* 24, 5374.—The heart of *Tropidonotus natrix* or *Vipera berus* is only $1/1000$ as sensitive to strophanthin as the frog heart. A qualitatively similar resistance was observed against the whole digitalis group, possibly due to decreased absorption by the cardiac muscle. The symptoms of poisoning, and the effects of cations, particularly Ca and K, are the same as in frog hearts. The stimulating and conducting portions of the snake heart are comparatively more sensitive to digitalis than the contracting portions. Saponin and saponin-like toxins affect snake and frog hearts similarly. The blood of *Tropidonotus natrix* and that of *Vipera berus* are not toxic for the heart of these snakes. II EAGLE

Treatment of morpheum by insulin and grape sugar. GUNTHER ANTON AND JOSEF JACOBI. *Alin. Wochschr.* 9, 1547 (1930). II EAGLE

The differentiation of the total blood sugar into glucose and galactose following the administration of galactose. II PETOW, II KOSTERLITZ AND II N. NAUMANN. *Klin. Wochschr.* 9, 1549-50 (1930).—In normal patients, there is little change in the blood glucose following the administration of galactose. The latter disappears from the blood within 2 hrs., reaching a max. of 0.02-0.07% within the first hr. In cases with hepatic damage the blood galactose may rise as high as 0.14%, and is still demonstrable after 2 hrs. In diabetes the galactose curve is normal but there is a surprising increase of blood glucose which may double its original value. II E

Effect of insulin upon secretion. PEYER F. MEYER. *Alin. Wochschr.* 9, 1578-81 (1930).—Insulin increases the acidity and usually the quantity of gastric juice within 15-90 min. The blood sugar change is the inverse of this change, reaching min. values when the gastric secretion is at its max. Diabetics who do not become hypoglycemic following insulin show no gastric reaction. In achylia, insulin like histamine, is ineffective. II EAGLE

The retention of uroselectan in the human body. W. TOURNÉ AND F. DAMM. *Klin. Wochschr.* 9, 1581-3 (1930).—With normal kidney function there is no demonstrable uroselectan in the blood 4 hrs. after its administration; residual quantities of < 0.5 g. in the blood show partial impairment of kidney damage, > 0.5 g. after 4 hrs. indicates a high degree of kidney damage. II EAGLE

Effect of caffeine in hypoglycemia. LUDWIG POPPER AND SUSANNE JAHODA. *Klin. Wochschr.* 9, 1585-6 (1930).—The symptoms of insulin shock are not related to the hypoglycemia and can be relieved by caffeine, which does not affect the blood sugar. II EAGLE

The effect of the administration of bilirubin upon the blood. LUDWIG POPPER. *Alin. Wochschr.* 9, 1770 (1930).—The intravenous injection of 0.05 g. of bilirubin into humans causes a transitory (24 hrs.) increase in hemoglobin without affecting the red cell count. II EAGLE

Clinical experiences in the treatment of post-encephalitic phenomena with harmine. HEINZ FRANK AND OTTO SCHLESINGER. *Alin. Wochschr.* 9, 1854-6 (1930).—Good results are reported, increased by the simultaneous administration of scopolamine. II EAGLE

The vascular effect of minute iodine concentrations. GEORG BARKAN AND SALME PRICE. *Alin. Wochschr.* 9, 1872 (1930).—No definite effect is produced by I_2 in concns. between 1:50,000 and 1:10,000,000 upon the caliber of rabbit ear veins. II E

Rhusinic acid, an active constituent of *Rhus succedanea*, L. KOJIRO TERAUCHI. *Tôhoku J. Exptl. Med.* 16, 123-56 (1930).—The chem. and pharmacol. properties of an

acid isolated from the fruit of *Rhus succedanea* are described in detail. It is primarily cardiotonic, is sympathomimetic and is very toxic for rabbits (lethal dose = 0.0067 g per kg).

H. IAGUI

Effect of camphor upon the adrenaline output and the blood sugar content in non-anesthetized, non-fasted dogs. KOGORO OIKAWA AND TATSUNAKO INABA. *Tokoku J Exptl Med* 16, 204-310 (1920).—Camphor olive oil injected subcutaneously into dogs in quantities greater than 1.75 g per kg caused increased secretion of adrenaline, hyperglucemia, increased pulse rate and respiration, lasting for as long as 2 hours with the max effect in 30-60 min.

H. IAGUI

The effect of certain narcotics (urethans) on permeability of living cells to water. BALDWIN LUCKE. *Biol Bull* 60, 72-9 (1931).—Narcotics in the presence of sea water do not decrease permeability to water beyond the value normally found in the fluid. There is a tendency to reduce permeability to water. Narcotics of this type are less effective in this respect than are bivalent cations. This effect, too, depends on the chem compn of the medium in which the narcotizing compd is dissolved. F. G. G.

Action of ephedrine on the dog heart (electrocardiograph studies). L. E. CORLINO. *Compt rend soc biol* 92, 1525-7 (1928). *Physiol Abstracts* 14, 521.—Synthetic ephedrine-ephedromine (Merck) produces certain modifications in the electrocardiogram of the dog. In small doses it causes bradycardia, sinusoidal and auriculo-ventricular block, ventricular extra systole and increase of the T wave. In large doses it induces tachycardia and affects the whole conducting system including Purkinje's fibers. Finally the heart fails in fibrillation.

G. G.

Intra-ocular tension and physico-chemical properties of the vitreous humor. T. REDLON AND P. REISS. *Compt rend soc biol* 92, 1550-8 (1928). *Physiol Abstracts* 14, 531. cf. C. A. 22, 3122.—Swelling of the vitreous humor depends upon its retraction. Isotonic solns of HCl and H_2PO_4 were injected into the vitreous humor of rabbits. The tension fell in 30 min, and the normal condition was restored in 30 days. All injections gave less characteristic results. Hypotension often follows the rise. Injection of gelatin causes a marked reduction lasting about 8 days. H_2PO_4 was injected in a case of glaucoma and caused a marked lowering of tension with relief which lasted a month. The effect of gelatin is due to its action on the isoelec point of the proteins of the vitreous humor.

G. G.

Action of mercurial diuretics on hydremla, chloremia, azotemia and the urinary elimination. I. HATZIGANU, I. GABRILA AND DORNIC. *Compt rend soc biol* 92, 1813-4 (1928). *Physiol Abstracts* 14, 258-9.—Neptal, solytargin and novasural given to man on a fixed diet cause identical variations in the blood and urine of subjects with different cardiac and renal affections. Hydremla is increased, reaching a max 5 hrs after injection. The Cl of blood is increased but not to the same extent in all cases. The N of the blood is practically unchanged. Intense diuresis occurs on the day of injection. The output of Cl on the day of injection is 2 to 7 times greater than on the day preceding. Urea is only slightly increased but not in all cases.

G. G.

Action of insulin and thyroxine on endocrine glands. J. VATRIN AND P. FLORENTIN. *Compt rend soc biol* 100, 111-3 (1929). *Physiol Abstracts* 14, 581.—After injection of insulin the thyroid becomes congested, and colloid degeneration increases in the hypophysis. Thyroxine in increasing doses causes a marked reaction in the pancreas, the islets of Langerhans are multiplied and new islets are being formed at the expense of exocrine acini. Insulin causes hyperplasia and thyroxine degenerative effects.

G. G.

Ergotamine, yohimbine and post-hemorrhagic hyperglucemia. I. I. NITZESCU. *Compt rend soc biol* 100, 386-8 (1929). *Physiol Abstracts* 14, 448.—Ergotamine and yohimbine inhibit the hyperglucemia produced by adrenaline. Hemorrhage causes hyperglucemia with increase of P and Ca. The latter form by hyperglucemia is also inhibited by the drugs mentioned, and it is therefore not due solely to increased glycogenolysis following increased output of adrenaline.

G. G.

Action of camphor, hexetone and sodium salicylate on cestodes and *Ankylostoma* of the dog. S. F. GOMES DA COSTA. *Compt rend soc biol* 100, 600-1 (1929). *Physiol Abstracts* 14, 671.—d Camphor in 3 M/1000 soln is inactive on the *Tenia serrata* of the dog. Camphor has a stronger effect. Hexetone in 5 M/1000 has a strong effect, and causes paralysis in twice this concn. Part of this action is due to the salicylate in which hexetone is dissolved. The action of these reagents on cestodes is somewhat less than on *Ascaris*. On the whole, the action of the camphors on *Ankylostoma* is practically nil. Hexetone is less active on *Ankylostoma* than on *Tenia*.

G. G.

Effect of sodium carbonate on polycythemia after muscular exercise. I. I. NITZESCU AND O. MIHAILESCU. *Compt rend soc biol* 100, 680-9 (1929); *Physiol Abstracts*

14, 4%—Among the metabolites of muscular exercise supposed to stimulate splenic contraction lactic acid occupies a prominent place. Na_2CO_3 injected intravenously opposes this action, which is therefore apparently due to the acid. Psychic excitation and lack of oxyhemoglobin are probably secondary agents. It is important to note that, while the alkali prevents splenic contraction and polycythemia, it does not prevent hyperpnea after exercise. G G

Action of chloroform on mammalian plasma and serum. P BORDET *Compt rend soc biol* 100, 751-3(1929), *Physiol Abstracts* 14, 283-4. cf *C A* 24, 891—If oxalated rabbit plasma is coagulated under the influence of CHCl_3 after several hrs only traces of thrombin are found in it. CHCl_3 coagulation of ordinary plasma yields abundance of thrombin. CHCl_3 does not excite the formation of thrombin in the absence of Ca. A serum resulting from normal coagulation rapidly loses its thrombin whereas that arising in presence of CHCl_3 maintains activity over long periods. G G

Chlorine, bromine and iodine and cardiac excitability. L DE BORGRAAF *Compt rend soc biol* 101, 167-8(1929), *Physiol Abstracts* 14, 350—Br can completely replace Cl as an anion in the perfusing fluid of the frog heart without causing any important changes in the beat. The irritability, however, is diminished. Iodine used to replace Cl entirely causes distinct changes. The heart is slowed and the irritability diminished. G G

Action of sodium glycocholate and sodium taurocholate in Ringer-Locke solution on isolated frog heart. D MANTIP AND S CIOARCEIU *Compt rend soc biol* 101, 225-6(1929), *Physiol Abstracts* 14, 375—Bradycardia was never observed as a result of the action of a weak soln of bile salts on the frog heart. Tachycardia was common, except in case of concentrated soln which caused, besides bradycardia, intoxication of the myocardium and tendency to arrest the beat. G G

The pancreatic function in insulin hypoglycemia. B A HOUSSAY, J T LEWIS AND V G FOGLIA *Compt rend soc biol* 101, 239-41(1929), *Physiol Abstracts* 14, 380—After the injection of insulin into (a) normal dogs, (b) dogs without pancreas and (c) dogs with cervical grafted pancreas the glucemic fall is similar in all, but the restoration is slowest in the last mentioned animals. This fact may be due to the suppression of the extrinsic nervous system, which influences the pancreatic endocrine function and which is stimulated by the hypoglycemia. G G

Effect of continued injections of insulin in depancreatized dogs. B A HOUSSAY, J T LEWIS AND V G FOGLIA *Compt rend soc biol* 101, 241-5(1929), *Physiol Abstracts* 14, 380—There is a basal secretion of insulin which increases or diminishes according to the glucemic fluctuation. The amt. can be extd by injections of insulin into depancreatized dogs. It is necessary in such cases to inject a regular dose of 0.01 unit per kg per hr. G G

Injection of sodium nucleate in dementia precox. H CLAUDE, P SCHIFF AND A DIMOLESCO *Compt rend soc biol* 101, 539(1929), *Physiol Abstracts* 14, 672—A markedly unresponsive condition of the nervous system to injection of nucleate of soda and to other kinds of interference is generally recognized. Yet in cases which show no improvement a profound reaction of the organism is indicated by certain changes in the blood. Leucocytosis is an early occurrence, chiefly affecting mononuclears. It may be exhibited before any signs of fever set in. eosinophils are increased later. The discordance between phys and psychic reactions is very marked in dementia precox. G G

Influence of the thyroid on reticulo-endothelial blocking. P GOEBEL *Compt rend soc biol* 101, 596-7(1929), *Physiol Abstracts* 14, 658—Injection of colloidal Ag which blocks the reticulo-endothelial system and causes hyperglucemia in normal dogs does not cause hyperglucemia after removal of the thyroid. The hyperglucemia is a function of the glycogen reserve and it may fail to occur when colloidal Ag is injected into a normal dog after thyroid feeding. Blocking of the reticulo-endothelial system causes no change in the cholesterol content of the blood in the dog after thyroidectomy. G G

Secretion and glucemia. E ZUNZ *Compt rend soc biol* 102, 339-40(1929), *Physiol Abstracts* 14, 576, cf *C A* 24, 413—Secretin free from hypotensive properties lowers glucemia in the dog, and the degree of this effect is in proportion to that on pancreatic activity. But preps. of secretin have been found which produce the effect on the pancreatic juice without hypoglycemia. Thus, the two actions are not caused by the same chem compd. G G

Histamine and alkalosis. R RAFFLIN AND P SARADJICHVILI *Compt rend soc biol* 102, 559(1929), *Physiol Abstracts* 15, 39—Subcutaneous injections of histamine

in epileptics and alcoholics produced a condition of alkalosis due to gastric and urinary secretion of Cl. A max. of alk. was reached 20 min. after the injection. The max. of Cl output by the kidney occurred in 10 min., then there was a fall due to gastric secretion followed by a 2nd max. 27 min. after the injection. G. G.

Antagonism of the base tropine (tropanol) and pilocarpine on the submaxillary gland. P. HAZARD. *Compt. rend. soc. biol.* 102, 574-5 (1929). *Physiol. Abstracts* 15, 19, cf. C. A. 24, 891. Tropanol checks the secret on of saliva evoked by injection of pilocarpine and pilocarpine increases the secretion when inhibited by tropanol. G. G.

Changes in toxicity of uranyl nitrate in the rabbit. M. GARTNER AND J. MAYER. *Compt. rend. soc. biol.* 102, 666-671 (1929). *Physiol. Abstracts* 15, 171. The authors formerly stated that 1 Gm. $UO_2(NO_3)_2$ injected subcutaneously in the rabbit was always fatal in a dose of 0.003 g. per kg. From a series of experiments they show now that the fatal dose varies under certain circumstances between 0.001 and 0.004 g. per kg. A cause of this variation apart from the breed, is the sex of the animal and the season. Females are more resistant than males and rabbits injected in the summer are less resistant than in cooler weather. G. G.

Action of ephedrine on the apical and venous pressures of the dog before and after yohimbization. M. LOEWER, A. LEWISER AND J. PATEL. *Compt. rend. soc. biol.* 102, 690-6 (1929). *Physiol. Abstracts* 15, 191-2. In the chloralosed dog the intravenous injection of 5 mg. of chloride of ephedrine causes a unchanged rise of blood pressure in the carotid artery. In the yohimbized dog the carotid rise of pressure is suppressed. Ephedrine acts on the meningeal pressure independently of the arterial system. Its action is the opposite of that of adrenaline which augments the meningeal pressure even after yohimbization. G. G.

Cholagog action of naptal. P. CHARBENT, P. CHARBENTAT, M. MAXIMIN, J. PASTIN AND HENRIETTE. *Compt. rend. soc. biol.* 102, 911-2 (1929). *Physiol. Abstracts* 15, 199, cf. C. A. 25, 352.—Increasing the mol. wt. of cholagog aromatic acids by substitution increases their effect on the flow of bile. The introduction of 11g into the formula of Na salicylate is effected in naptal. In doses of 2 to 4 cc. it doubles the amt. of bile, in 1 ccpt the bile was quadrupled. The 11g per se does not affect the action. There is a certain parallelism between the cholagog and diuretic action of naptal. A diuretic like $CaCl_2$ on the other hand, has an inhibitory action on the secretion of bile. G. G.

Influence of the non-saponifiable substances of ox spleen on anemia of the rabbit. C. SOLTA, J. TAVIN AND J. PAU. *Compt. rend. soc. biol.* 102, 1022-3 (1929). *Physiol. Abstracts* 15, 151.—Spleen rabbits require 10 to 20 days for the regeneration of the erythrocytes. A rabbit which receives 1 g. of cholesterol daily requires 14 days for this result. Given a daily dose of 1 g. of non-saponifiable substance (X), the rabbit needs but 7 days for the regeneration. G. G.

Study of β -aminobenzoate of *N*-diethyleucinal (panthesin). A. PAIN. *Compt. rend. soc. biol.* 102, 1043-4 (1929). *Physiol. Abstracts* 15, 192.—The formula of panthesin has the same aromatic nucleus as novocaine, $C_{11}H_{11}COOCH_2CH_2N(C_2H_5)_2$. It may be sterilized by boiling. It does not inhibit the action of adrenaline on the rabbit ear. Anesthesia of the cornea with a 2% soln. of panthesin lasts 22 min., as against 27 for cocaine. The relative toxicity of panthesin, i. e., the ratio of toxicity to anesthetic power, is 2 or 3 times less than that of novocaine. The a. v. toxicity of panthesin is greater. G. G.

Cholagog action of certain phenolic acids. M. MAXIMIN AND A. HENRIETTE. *Compt. rend. soc. biol.* 103, 3-5 (1930). *Physiol. Abstracts* 15, 195-9.—Sodium benzoate is inactive toward biliary secretion in the dog in doses of 1 g. per 20 kg. The salicylate gives inconsistent results. The *m*-hydroxybenzoate has a strong cholagog action. The diphenolic acids, catecholic and caffeic, greatly increase the output of bile with increased wt. of solvent. *p*-Hydroxybenzoate is practically without effect. It seems that the position of the *m*-hydroxyl group favors the development of cholagog properties. Vanillic acid, in which the *OH* in the meta position is replaced by OCH_3 , has a very marked cholagog action. Though the phenol function intervenes in the effectiveness of the compd., it is not in itself a sufficient factor. The role of the carboxyl group, as previously described, is further incerted on. G. G.

Action of insulin in tissue culture. H. PEARSON AND J. ROSENBLUM. *Compt. rend. soc. biol.* 103, 50-2 (1930). *Physiol. Abstracts* 15, 173.—Insulin acting directly on mesenchymatous cells in artificial culture is without influence on the formation of glycogen, and, far from increasing the consumption of sugar, it diminishes this consumption. The diminution runs parallel with a disturbing influence on growth *in vitro*. G. G.

Action of dinitro- α -naphthol on the pigeon. P. V. UYTVANCK *Compt rend soc. biol* 103, 29-31(1930), *Physiol Abstracts* 15, 103 —The pigeon after injection of dinitro- α -naphthol exhibits polypnea; 12 min later the rectal temp. rises and reaches 46.5°. In 18 min peripheral vasodilatation, hyperglucemia and diminution of muscular and hepatic glycogen accompany the hyperthermia. The phosphates of blood and muscle are increased. The Ca content of blood does not change. The hyperthermic effects of dinitro- α -naphthol are not suppressed by exposure of the reagent to the rays of a Hg vapor lamp. G G

Action of calcium and potassium on the isolated heart. R. KIRSCH AND H. FRÉDÉRICQ *Compt rend soc biol* 103, 34-6(1930), *Physiol Abstracts* 15, 157 —The chronaxie of the isolated heart of the tortoise is lengthened by an excess of Ca or a deficiency of K. It is diminished in the reverse conditions. If the perfusion with the modified Ringer soln contg excess Ca or deficient K be prolonged, then the chronaxie slowly returns to its original normal value. A fresh perfusion now with normal Ringer soln causes a new variation of the chronaxie in the opposite direction to the first change. G G

Karyokinetic action of arsenious acid and alkalosis. M. ROCMANS *Compt rend soc biol* 103, 42-3(1930), *Physiol Abstracts* 15, 193-4 —In the guinea pig and dog increased multiplication of the cells in the lymphoid organs occasioned by H_2AsO_3 runs concurrently with a wave of alkalosis. The soly of the Ca ion is diminished, and this effect is proportional to the dose of H_2AsO_3 . The results obtained with H_2AsO_3 correspond with those obtained in peptone shock. The curves of cellular division and alkalosis run parallel. In the dog the normal pH of 7.42 rises to 7.52, and that of the guinea pig rises from 7.29 to 7.42. The acid base regulating mechanism does not alter, for the variations of $NaHCO_3$ and H_2CO_3 are mutually balanced. G G

Arterial tension and leucocytes. Action of acetylcholine. L. JUNG AND P. COLLET *Compt rend soc biol* 103, 79-81(1930), *Physiol Abstracts* 15, 158 —The lowering of arterial pressure caused by injection of acetylcholine in dogs is accompanied by leucopenia. A return to a normal leucocyte count accompanies the restoration to the normal of arterial pressure, but it is slower than the rise of pressure. G G

Latent period of fluoride intoxication. H. CRISTIANI *Compt rend soc biol* 103, 202-4(1930), *Physiol Abstracts* 15, 126 —The bones of animals from districts where there are no factories emitting F contain very small quantities of this element in their compn, less than 0.05%. The bones of animals from the neighborhood of Al works contain 0.379%. Therefore, in regions where F is present, even in min quantity, there is a condition of latent fluorosis in the cattle. G G

Strychnine-soap complexes. L. VELLUZ *Compt rend soc biol* 103, 302-3(1930), *Physiol Abstracts* 15, 124 —Soaps tend to neutralize the toxicity of alkaloids but to a less extent than that of toxins. A mixture of oleate of Na and strychnine, boiled for 15 min on a water bath, may be injected into a guinea pig in 4 times the fatal dose of strychnine without harmful effect. Increase of the amt leads to the typical symptoms. Repeated injections of the complex during a period of 2 months did not produce any immunity either to pure strychnine or to the complex, and it is obvious that immunization against crystalloid poisons cannot be effected by conferring colloidal characters on them. G G

Suprarenal capsules in hyperglucemia produced by decamethylenediguanidine. X. CHAIKOVITCH *Compt rend soc biol* 103, 328-30(1930), *Physiol Abstracts* 15, 117 —Synthalin produces hyperglucemia in the rabbit by its action on the suprarenal capsules. Bilateral extirpation of the capsules prevents the hyperglucemia following injection of synthalin, extirpation of both suprarenal glands does not *per se* cause variations in the amt. of free sugar in the blood. G G

Vascular effects of acotine. H. FIKERFORS *Compt rend soc biol* 103, 443-5(1930), *Physiol Abstracts* 15, 124 —Acotine exerts a double action on blood vessels—constrictor in small doses and vasodilator in large doses. The latter is the more common effect, the first being sometimes doubtful. Atropine does not inhibit the action. Acotine inhibits the constrictor action of $BaCl_2$ and therefore may be supposed to act on the muscle. G G

Alteration of the thyroid gland in fluoride intoxication. H. CRISTIANI *Compt rend soc biol* 103, 554-6(1930), *Physiol Abstracts* 15, 193 —Guinea pigs were given fluorides in small doses over long periods and their thyroid glands compared with those of controls. In the glands of the former, lesions were common and marked. The animals died after an av duration of 53 days. G G

Macroscopic alterations of the hypophysis in fluoride intoxication. H. CRISTIANI *Compt rend soc. biol.* 103, 556-7(1930), *Physiol Abstracts* 15, 193 —Chronic poisoning

with fluorides causes diminution of vol of the hypophysis, which becomes flat in appearance. The thickness in the guinea pig is reduced from about 2.25 to 1.7 mm. G G

Immediate hyperglucemia with use of insulin. F RATHERY, R. KOURILSKY AND (Mlle) Y LAURENT *Compt rend soc biol* 103, 563-4(1930), *Physiol Abstracts* 15, 173—Injection of insulin in the normal dog causes an immediate rise of the blood sugar, which is transient it is followed quickly by a fall. This hyperglucemia is not const., but it is frequent. The phenomenon seems to be independent of the nature of the insulin employed, the dose and the path of injection. G G

Insulin hyperglucemia in different vascular regions. F RATHERY, R. KOURILSKY AND (Mlle) Y LAURENT *Compt rend soc biol* 103, 565-6(1930), *Physiol Abstracts* 15, 173—The hyperglucemia immediately following the injection of insulin in the normal dog is more marked in certain parts than in others. It is especially pronounced in the hepatic circulation. There is increased discharge of sugar into the hepatic vein and a notable increase of portal blood sugar. G G

Cardiovascular effects of choline derivatives administered by the digestive tract. M. VILLARIT, L. JUSTIN BÉSANÇON AND R. CACHIRA *Compt rend soc biol* 103, 583-4(1930) *Physiol Abstracts* 15, 190—Methylacetylcholine introduced directly into an exposed loop of intestine of the dog causes a fall of blood pressure which is maximal in 1 min. Absorption is extremely rapid. The arterial fall of pressure is very prolonged. The heart is scarcely affected, the fall being due to dilatation of the arterioles. Bromocholine used in the same way causes a fall of pressure which is less marked, but bradycardia sets in with respiratory trouble. The animal dies in 50 min. from a dose of 60 mg. per kg. G G

Utilization of pentoses by the animal organism. Action of intravenous injection of pentose on lactose secretion. I. I. NITZESCU AND M. BENETATO *Compt rend soc biol* 103, 1119-22(1930), *Physiol Abstracts* 15, 119-20—After the intravenous injection of pentose the amt of sugar in the milk is increased. Injection of sucrose causes no change in the output of lactose. The lactose forming power of rhamnose is not as great as that of glucose, but it is equal to that of galactose. G G

Acute mercury poisoning, its treatment and alterations in chlorine, water, nitrogen and acid-base relations during its course. A. LANDAU, J. GLASS AND I. BEILES. *Wiener Arch inn Med* 20, 319-42(1930)—Acute Hg poisoning with Hg oxycyanate was treated by Bi therapy and by administration, after the period of anuria had passed, of NaCl by mouth, and rectal and subcutaneous injections of physiol NaCl soln. A chem. study was made of the blood and urine during the 5 definite stages of the poisoning. During the first period of anuria there was a retention of water without edema, a marked chloropenia of the plasma and blood corpuscles, a constantly increasing azotemia and a lowering of the alkali reserve, forming a complex which is typical for severe Hg poisoning. The chloropenia is probably due to loss of Cl in vomiting and diarrhea, as in mild cases without these symptoms the chloremia remains normal. The azotemia is very marked, it depends in part on the anuria and on a toxic breaking down of the cell proteins. There is a retention of urea, uric acid, creatinine and other N-contg compds. The azotemia continued to increase for a time after the administration of NaCl had begun. Administration of NaCl at later stages did not increase the azotemia. Usually in acidosis the Cl index of the blood corpuscles rises from increase in Cl content in the corpuscles. Here, however, there is a marked decrease in alkali reserve without a shifting of the NaCl from plasma to blood corpuscles. In spite of a decreased alkali reserve the urine remained alk. Hyperventilation of the lung from toxic stimulation of the respiratory center by Hg or from toxic split products from cellular proteins, and the lowering of the CO₂ tension in the alveolar air bring about a decrease in alkali reserve of the plasma with excretion of basic substances by the kidneys. HARRIET F. HOLMES

Lymphadenoma produced by injection of coal tar. L. MERCIER. *Compt. rend. soc. biol.* 103, 125-6(1930), *Physiol Abstracts* 15, 198—Injection of coal tar in olive oil into the peritoneal cavity of the mouse caused generalized hypertrophy of the lymphatic glands, and splenic and pulmonary nodules. In some of the descendants of an animal thus treated lymphadenoma appeared spontaneously. Injection of coal tar into the progeny of sensitized animals caused a rapid appearance of lymphoid growths with an increase in the no. of pos. results. E. C. K.

Reaction of thyroid gland to hypophyseal extract. M. ARON. *Compt rend. soc. biol.* 103, 145-7(1930), *Physiol. Abstracts* 15, 180.—Details are given of the histological changes which occur in the thyroid gland as a result of hypophyseal injection in the guinea pig. The vesicles become empty and the epithelial cells are enlarged. The

evidence of marked activity is more easily observed and less equivocal in the thyroid than in any other gland F. C. K.

Blood-vessel wall changes produced by vigantol and their regression in animals. ALICE SCHIFF. *Arch Path Anat (Virchow's)* 278, 62-83(1930)—Feeding vigantol (com. irradiated ergosterol) to cats and rabbits in daily doses of 2-4 mg. for 1-3 months resulted in changes in the media of the arteries. The changes included degeneration, proliferation and calcification. Regression of the lesions could not be definitely established, on the other hand, there was some evidence of subsequent progression. E. R. LONG

Initial stages of vigantol injury in rabbits. ERNST LAAS. *Arch Path Anat (Virchow's)* 278, 316-61(1930)—Single doses of 30 mg. vigantol per kg. animal weight produced in rabbits the following changes: calcification in the media of arteries, beginning in the elastic fibers without a recognizable pre stage, and passing in the aorta from a localized nodular medial necrosis to a more diffuse degeneration, secondly, calcification in various organs, dependent in turn upon degeneration caused by injury to the vessels concerned in the local blood supply. E. R. LONG

Sectioning of the splanchnic nerve and the effect of cholic acid on the creatinine excretion. AIIJAO TAKU. *J. Biochem (Japan)* 12, 371-81(1930)—Subcutaneous injection of cholic acid causes a diminution in the creatinine excretion of the rabbit, but this effect disappears after sectioning of the splanchnic nerve. The hypoglycemic influence of cholic acid is likewise removed by sectioning of the splanchnic. All of this indicates that cholic acid exerts its action not peripherally but centrally. S. MORICULIS

Influence of hyperglucemia and of extirpation of the adrenal bodies on the amylolytic power of saliva and blood. TULLIO GAYDA. *Arch ital biol* 84, 55-66(1930)—See *C. A.* 25, 328, 338. S. MORICULIS

Behavior and action of organic sulfur compounds in the organism of the dog. I. Action and fate of thiophene in the metabolism of the dog. ANAST. A. CHRISTOMANOS. *Biochem. Z.* 229, 248-51(1930)—Thiophene is detd. colorimetrically as follows: To each of 2 test tubes are added 5 cc. concd. H_2SO_4 and 2 drops of satd. $CuSO_4$ soln., then 0.05 cc. *N* lactic acid. This is thoroughly mixed and heated in boiling water exactly 3 min. After cooling to about 10° a known amt of thiophene in alc. soln. is added to one and the unknown soln. to the other test tube, this is mixed and left at room temp. for 10 min. when a color from rose to deep red develops, depending upon the concn. of the thiophene. The colors should be matched quickly in a colorimeter, and the results calcd. by the formula $C = C_1 + S_1/S$, where C_1 = concn. of known soln., S_1 and S are the colorimeter readings of the known and unknown solns. As little as 0.013 mg. still gives a recognizable color. When fed to dogs 5.7-12.2% of the thiophene are excreted in the urine as a conjugated product and part is assumed to change to CH_2 , CH_2SH in the intermediate metabolism. Thiophene does not depress the total *N* metabolism. Thiophene causes a characteristic ataxy reaction in the dog. S. M.

Influence of different foods on the amount of unsaturated fatty acids with four double bonds in the blood. N. BEREVO. *Biochem. Z.* 229, 323-8(1930)—The amt of fatty acids with 4 double bonds in blood increases, after a preliminary drop, following the ingestion of fat. This preliminary decrease does not occur if carbohydrate is fed together with the fat. Carbohydrate or protein alone has no effect. S. M.

Assimilability and toxicity of racemic lactic acid. OTTO FÜRTH AND PAUL ENGEL. *Biochem. Z.* 229, 381-96(1930)—The ingestion of fermentation lactic acid causes in man the loss of 20-30% in the urine. The lactic acid is detd. in human urine by pptg. the urine with phosphotungstic acid and removing from the filtrate the glucide like substances by pptn. with $Ca(OH)_2$ and $CuSO_4$. In rabbit urine lactic acid cannot be so detd., the long and difficult preliminary extn. with ether in a Lindt liquid extractor is necessary. For the rabbit lactic acid administered orally proved much more toxic than previously supposed. Although subcutaneously as much as 2.8-3.6 g. per kg. can be injected without any obvious harm rabbits died when fed 1.6 g. and even as low as 0.6 g. lactic acid per kg., also when that quantity was given over a period of 3 days. For mice the lethal dose for subcutaneous injections of the *Na* lactate was 2-4 g. per kg. Rats, on the contrary, can withstand prolonged administration of 2 g. or more per kg. without any ill effects. The high toxicity of the ingested *L*-lactic acid for rabbits and its poor assimilability also in man throw doubt upon its value as a food substance. The use of small quantities of lactic acid in lemonades, fruit juices or confectionary is probably harmless. Alkalosis induced by $NaHCO_3$ did not alter the effect of lactic acid in rabbits and mice. Nor was the utilization of the lactic acid by the rabbit materially altered by adrenaline, phlorhizin or loading with sugar. Mild

P Intoxication caused in rabbits a large loss of lactic acid, followed, however, by increased retention when the intoxication passed off. *In vitro* expts on the catalytic oxidation of lactic acid on C surfaces fail to reveal any difference between the d- and l acid, as is observed in physiol expts. S. MORGULIS

Does adrenaline exert a direct action on the base of the mid-brain? R. H. KAHN AND C. RINDT. *Endokrinologie* 8, 1-4 (1931).—Following the application of a 1:1000 adrenaline soln to the surface of the base of the mid brain there is no evidence of a sympathetic stimulation of end organs or of any injury. This indicates that the direct application of adrenaline to nerve tissue is without effect. On the contrary, a strong rise in blood pressure results promptly from the laradic stimulation of the base of the mid-brain. S. MORGULIS

Effect of alkaline extracts of the anterior lobe of the hypophysis on the genitalia of female rats. M. REISS, H. SELVE AND J. BALINT. *Endokrinologie* 8, 15-22 (1931).—Alk. exts from the anterior hypophysis exert an inhibiting effect on the hormone prepn from the urines of pregnant women. This inhibiting action was especially manifested on the epithelium of the uterus. The striking changes produced by the injection of the sex hormone were absent if this was accompanied by an injection of the alk. ext. The uterus hypertrophy does not occur in animals treated with both exts and the great hypertrophy of the ovary likewise fails to develop. S. MORGULIS

Studies on relation of chemical constitution and biological action: amyostatic poisons. TAKAOKI SASAKI AND ICHIRO OTSUKA. *J. Biochem. (Japan)* 12, 429-58 (1930).—Synthetic products made from aminophenylalanine with glycine anhydride possess very interesting biol. actions. These compds., which are the HCl salts of 3-aminohydrocarboxystyrls (3-Asty), act specifically upon the brain stem and cause, in rabbits given subcutaneous injections, disorganization of the myostatic and statokinetic responses resembling the decerebrate rigidity or the alimetic-rigidity syndrome observed in human pathology. The amyostatic action of local anesthetics like novocaine, tetracaine and cocaine is similar to that of the 3-Asty and its derivs., which also seem to have local anesthesia effects. The original must be consulted for a detailed account of the chem. prepn of these various astyl compds and their biol. effects. S. MORGULIS

The cause of the diminished effect of the vagus on the heart under insulin. C. I. RAHL AND P. MALM. *Stand. Arch. Physiol.* 61, 64-70 (1931).—The diminished chronotropic effect of vagus stimulation upon the heart under the influence of insulin is attributable to a lowered sensitivity of the heart. There is some evidence to show that the depressing effect of the insulin on the response to vagus stimulation is associated with an increased K content of the heart. S. MORGULIS

Copper studies. C. G. SANTESSON. *Stand. Arch. Physiol.* 61, 70-112 (1931).—Liver injury from Cu intoxication may arise in a variety of ways. In some instances of acute intoxication or intoxication for a long period with very large doses the blood is the seat of primary injury. Icterus with hematuria and hemoglobinuria, even hemorrhagic nephritis follow, and sometimes there is also evidence of methemoglobin formation. The disruption of red cells occasionally produces thrombosis of the liver blood capillaries which lead secondarily to necrosis of the parenchyma. However, the liver injury may also result from an altogether different cause. In the author's expts with CuSO_4 , the primary change in the liver was a strong capillary hyperemia with extensive bleeding which disrupted the liver parenchyma. The clotting of the blood takes place readily, leading to capillary thrombosis and thus to necrosis without primary blood alterations. The hyperemia is due to dilatation of the finest blood vessels and is common in all metal intoxication (Hg, Bi, Cu and especially As). The bleeding, at least in the case of As and P results from injury of the vessel wall. S. never found icterus or cirrhosis of the liver, and thinks that possibly rabbits are not subject to icterus. A round cell infiltration was observed which might possibly have been an early stage of cirrhosis. The hyperemia from Cu intoxication extends also to the lung and frequently causes hemorrhage or pneumonia. Although intestinal changes have frequently been described, S. found little evidence in his expts. Likewise, the kidney was only slightly affected, although occasionally a little albumin appeared in the urine but never any bde pigment. Generally the intoxication with Cu or Bi has similar effects, but the former acts more upon the liver, whereas the latter acts on the large intestine and cecum and on the kidneys. S. MORGULIS

Toxic action of dictamnina. A. OGATA. *J. Pharm. Soc. Japan* 50, 1124-33 (1930).—Thoms (C. A. 17, 2583) previously isolated dictamnina (D) from *Dictamnus albus* and studied its toxic action. O. repeated the study of the toxic action of D and found a slightly different toxic action. F. I. NAKAMURA

The action of formaldehyde on neuro-muscular excitability. D. BENNATI AND E.

HEARFIELD *Compt rend.* 190, 1522-5(1930) —The injection of 0.1 cc of 2% CH_2O in saline into the lymph sac of a 35 g frog produced no important changes in the ebronaxia or rheobase of the sciatic nerve gastrocnemius. The injection of 0.5 cc of the same soln caused loss of excitability apart from an appreciable modification of the 2 characteristics of this excitability. A. protocols are included. T. H. RIDER

Physiological action of aminomethyl-(3,4-dihydroxyphenyl)-carbinol. RAYMOND-HAMET *Compt rend* 191, 869-71(1930), cf Barger and Dale *C A* 5, 109 —Aminomethyl (3,4-dihydroxyphenyl)-carbinol (artirénol) (I) when injected prior to *d,l* adrenaline HCl (II) causes a greater hypertensive effect than the later injection of an equiv dose of II. However, when II is injected it produces a greater hypertensive effect than a later injection of I. Loup (*C A* 7, 4016) has shown that the hypertensive effect diminishes with repeated equal doses of II or other like substances. The action of I proves a little greater than that of II; therefore, the mechanism of the action must be different. In a dog with a sufficient prior dose of yohimbine (III), to reverse the action of II, I still causes hypertension, accompanied, however, by renal dilation, whereas I administered without a prior dose of III causes renal constriction. The action of I is, therefore, similar to that of ephedrine (IV) (cf R-H *C A* 22, 271). However, while strong doses of IV are hypotensive, fatal doses of I are still hypertensive. Similar to IV and contrary to II, the effects of I are not augmented by cocaine. T. H. R.

The chemistry of anesthesia. WILDER D BANCROFT AND GEORGE H RICHTEK *J. Phys. Chem* 35, 215-68(1931) —All types of anesthesia are considered explained satisfactorily by the theories of Binz and of Claude Bernard, that the reversible coagulation of the colloids of the sensory nerves produces anesthesia. The material offered by this paper, however, is limited to a discussion of general anesthesia, local anesthesia not being specifically included. A cursory criticism of other theories is followed by the development of the Binz Bernard theory, and a discussion of the 3 basic criticisms directed against it, namely (a) The concns. at which narcotics are active are much less than that required to flocculate the cell colloids. (b) Coagulation of the cell colloids represents the toxic effect of a narcotic and is irreversible, whereas narcosis is a reversible phenomenon. (c) The cases in which the dispersion of the colloids is decreased by narcotics are explained as being only apparent. In answer to (a) B and R call attention to the fact that the presence of electrolytes makes possible the flocculation of protein sols with a low concn of EtOH , and the assumption that the colloidal systems of the cell may be so sensitized in life that a small quantity of a narcotic will flocculate them. (b) and (c) are answered by the fact that with sensitized protein sols reversible flocculation with low concns of narcotics was accomplished *in vitro*, and that reversible flocculation was demonstrated in yeast cells. Facts which cannot be explained by other theories but which can be explained by the Binz Bernard theory are stressed. The stimulation of general protoplasmic irritability produced by low concns of narcotics is explained as dependent upon the decreased stability of the colloids caused by these low concns while narcosis is caused by higher concns. Attention is called to the fact that typical stimulants, such as caffeine and strychnine, behave as narcotics in high concn. The stimulating after-effect of narcotics is explained as due to the same colloidal state, passed through again in the process of peptization. Anesthesia caused by water, salts, heat, cold, or elec or mech means may also be explained as due to the effect on the colloidal state of protoplasm. The conclusion is reached that there is no relation between chem structure and physiol action on the grounds that structure merely det. the phys properties, which in turn, together with concn, govern the physiol action. All drugs are assumed to stimulate in some low concn and to depress in higher concns. Synergism is explained as follows. A small concn of a drug is adsorbed entirely upon the principal substrate. The addn of a synergistic drug which is adsorbed upon the same substrate displaces a certain amt of its predecessor which then affects the secondary substrates in the same manner as would a higher original concn. Antagonism would occur when the colloids of a cell are reversibly coagulated and the agent that caused the flocculation is displaced by a substance of weaker flocculation capacity for the given concn, in which case the bio-colloids will again be peptized by the electrolytes. NaBr and antipyrine which have a quieting effect on the nerves, produce a lowered irritability following upon increased dispersion of the colloids. Narcotizing agents are divided into 2 groups: direct narcotics, which coagulate the cell colloids by direct action (Et_2O , CHCl_3 , I OH) and indirect agents which interfere with some normal function of the cell, such as oxidation and the coagulation is produced by the accumulated waste products (NO_2 , C_2H_5 , HCN). T. H. RIDER

Investigation into the distribution of lead in the organism on basis of a photographic (radiochemical) method. SVEVD LOMMOLT. *J Pharmacol* 40, 235-45(1930) —A

soln of PbCl_2 contg a small amt of radium D was injected into mice and subsequently sections of the various organs, or the entire animal, were prepd and placed on photographic plates. When the plates were later developed the presence of the radioactive substance could be seen. Pictures showed that Pb is deposited in greatest amt in the liver, in the cortex of the kidney, and in osseous tissue. Very little was found in the brain. C. RIGEL.

The electrocardiogram of non-anesthetized dogs as modified by the intravenous injection of pitressin, atropine sulfate and vagus section. CHARLES M. GRUBER AND WM. B. KOENIG. *J. Pharmacol.* 40, 233-73 (1930). cf. C. 1 25, 745.—The following cardiac changes were noted: (1) a brief period of slowing, (2) a period of acceleration, (3) a prolonged period of slowing. During the latter phase the electrocardiogram showed shifting pacemaker, prolongation of auriculo-ventricular conduction time, sino-auricular block, partial auriculo-ventricular block, extra systoles, high T waves, high branching T waves and bigeminal pulse. After atropine or vagus section, the same changes were noted. C. R.

A study of the effect of morphine upon the respiratory center. A. H. MALONEY AND A. L. TATUM. *J. Pharmacol.* 40, 291-301 (1930).—Depression of the respiratory mechanism is more marked after neonal phenobarbital or (to a less degree) amytal and ethyl α -methylbutylbarbiturate acid, than after urethan, chloral hydrate or avertine. From observations on stimulation of the central end of the cutaneous nerve, stimulation of the intact skin and the effect of increased or decreased intrapulmonary pressure on respiration after morphine in rabbits it is concluded that morphine reduces the responsiveness of the center to acceleratory afferent nervous impulses and to CO_2 , but increases the effectiveness of inhibitory vagal impulses. C. R.

Iodine prophylaxis and endemic goiter. A. T. CAMERON. *Can. Pub. Health J.* 21, 495-500, 541-5 (1930).—An extensive discussion and review of the literature on the distribution of I in nature and its correlation with thyroid diseases, particularly simple endemic goiter. A close relationship between deficiency of I in the diet and simple goiter is indicated. Other possible causes of endemic goiter are discussed. There is strong evidence that in certain regions a water-borne infection is a factor, but this probably acts through an effect on I metabolism. It is concluded that iodized salt is the best medium for I prophylaxis. The I content of Canadian iodized salt (1 part KI or NaI to 10,000) seems unnecessarily high, and is perhaps 10-20 times greater than required. The evidence that iodized salt is potentially dangerous for adult (non-toxic) goitrous individuals is frequently open to criticism, and gravely exaggerated. Further study of this aspect is desirable. A bibliography of 60 references is appended. R. E. THOMPSON.

Influence of externally applied preparations to promote growth of hair. A. FORSTER. *Arch. exp. Path. Pharm.* 144, 363 (1929), *Schweiz. Apoth. Ztg.* 68, 545-7 (1930).—Results of systematic expts. on cats with 6 classes of medicaments (19) do not support the belief that the normal growth of hair can be influenced by the external application of supposedly sp. substances promoting growth, e. g., cholesterol, chem. degradation products of hair or "proteolytic enzymes." Aside from the effect of substances producing excessive stimulation of the skin, or of too high concns., treatment with each of the substances examd, even white vaseline, or 50% alc., although widely different in chem. and pharmacol. characters, may produce an increase in the growth of hair. This fact must be ascribed to the effect of mechanical stimulation, an increased flushing with blood, and perhaps also a cleansing of the skin. However the mechanical effect may be increased by the use of locally acting chem. stimulants. S. WALDBOTT.

Poisoning by seeds of the castor oil plant. P. LIPPAK. *Ber. Ungar. pharm. Ges.* 1928, No. 5, *Med. und pharm. Rundschau, Schweiz. Apoth. Ztg.* 68, 290-1 (1930), cf. C. A. 15, 2936, 17, 1844, 19, 2333.—An account is given of 3 cases of poisoning by ricinus seed, 2 cases were fatal after, resp. 2 1/2 and 5 seeds only had been eaten. S. WALDBOTT.

A study of the treatment of Sydenham chorea. B. WASMER. *Thesis*, Lyon, 1929; *J. pharm. Alsace Lorraine* 57, 285 (1930).—Arslyen has been found an efficient and innocuous remedy (cf. C. 1 22, 813). S. WALDBOTT.

Distribution of hydrastine between the erythrocytes and the plasma. R. FABRE. *J. pharm. chim.* [8], 12, 339-45 (1930), cf. C. A. 24, 4805.—Hydrastine-HCl was injected into a dog and 100-cc. samples of blood were taken after 1, 24, 48 and 72 hrs. Na citrate was added and plasma and erythrocytes were sepd by centrifuging. From each sample the alkaloid was extd. by the Stas-Otto-Oger method, converted into fluorescent hydrastine sulfate with H_2SO_4 and H_2O_2 and equal vols. of the solns. were photographed in Wood's light. Again, the erythrocytes proved to have a fixing power for

the alkaloid superior to that of the plasma Expts *in vitro*, however, showed no difference

Bee and wasp poison. DR VASCONCELLES *Rev Therap* 1930, No 9, *Therap Ber*; *Schweis Apoth - Ztg* 68, 501-3(1930) — Present knowledge of this poison and the methods of combating its effects on man are reviewed, cf Langer, 1826, *C A* 13, 2034, 14, 78, 15, 378 2127, 17, 2916, 3058, 18, 1531, 23, 900, etc Effective antidotes are an injection of 0.1% adrenaline soln (Samson), or of racedrine (racemic epinephrine) (Rubitschang) which prolongs the antidotal effect, or of ephrine which is a combination of both

S WALDBOTT

The chemical examination of *Sila cordifolia*, Linn (GROSU, DUTT) 17. Chlorotone as a preservative (POLAK) 17.

WOLFF-EISNER, A *Handbuch der experimentellen Therapie, Serum- und Chemotherapie: Ein Handbuch für die ärztliche Praxis und Klinik*. Munich J F. Lehmanns 236 pp M 18

I—ZOOLOGY

A A GORTNER

Insect oils and fats. J TIMON DAVID *Bull soc chim biol* 12, 395-403(1930), cf *C A* 23, 3513 — The larvae of more than 2 dozen insects studied contain quantities of oil varying from 0.94 to 24.0%. The 1 values of the oils vary from about 1.2 in the case of some Aphidians (these oils have a high sapon value) to 164.5 for *Salurnia pyri*, Sch. The non drying oils are most largely represented. The isolation of hexabromostearic acid by bromination of the acids from *Celaspidema atra*, Ol (m p 178.5°) confirms the presence of linolenic acid in the oil of this insect. Oleic acid is widely distributed in insect oils. The amt of unsaponifiable matter varies from 0.75% in the oil of *Ergates faber*, L. to 12.5% in that of *Heliosaurus menticornis*, Reitt. The oils of tropical insects have a lower 1 value than those of insects inhabiting cold countries. Females generally are richer in fat than males. Related species yield oils of a similar type

B C A

A study of the occurrence of trimethylamine in marine animals. A STANLEY COOK *Can Chem Met* 15, 22-3(1931) — Trimethylamine oxide was isolated from the muscle of various species of fish. A method is given for the quant estn of $(CH_3)_3NO$ in biol materials, based on the reduction to $(CH_3)_3NH$ and micro steam distn and titration of the liberated $(CH_3)_3NH$

W. H. BOYNTON

Occurrence of manganese in insects. II. A P VINOGRADOV AND M V. NEUSTRUBVA *Compt rend acad sci U R S S* 1930A, No 6, 127-32, cf *Compt rend acad Sci U R S S* 1929A, No 10, 227 — Mn was detd in a large no of insect species. The av % of Mn based on live wt was 10^{-4} , the extreme values being 2×10^{-4} and 5.2×10^{-3} . *Formica rufa* has the max content. The amt of Mn depends somewhat on the character of the food, is independent of sex and varies inversely with the atm of gaseous exchange in the insect, those always in motion having the smallest Mn content. With regard to distribution in the body, in *Formica rufa* the head contains an av % of 2.8×10^{-4} , middle portion 9×10^{-4} , posterior portion 3×10^{-4} based on the dry wt. The latter part of the insect contains the major portion of the alimentary tract and also the gland furnishing the poisonous secretion. Its high Mn content is not due to food residues in the tract since the food of these insects contains less Mn and more Fe than the whole insect. As Mn plays a catalytic role in some fermentations and as HCO_2H probably results from the action of an oxidative enzyme upon carbohydrate, the high Mn of the posterior part may be related to HCO_2H production

LEWIS W BUTZ

Occurrence of vanadium in marine organisms. A P VINOGRADOV *Compt rend acad Sci U R S S* 1930A, No 17, 465-7 — *Phallurus obliqua* and *Ciona intestinalis* contain V, *Malpadia affinis* and *Cucumaria frondosa* do not. It would seem that all Ascididae have a V-contg blood pigment. *Phallurus obliqua* contains 0.0302% of the live wt. The quant. method follows: Fuse the ash with Na_2CO_3 and KNO_3 , ppt Pb vanadate from the aq soln of the melt, dissolve the ppt and titrate the V. reduced by SO_2 , with $KMnO_4$

LEWIS W BUTZ

Molar concentration of snail blood, its factors and variations; influence of the state of activity of the animal. MARCEL DUVAL *Ann physiol physicochim biol* 6, 346-64(1930) — There is no definite relation between the state of activity of *Helix pomatia* and the molar concn of its blood as tested by the depression of the l p. NaCl constitutes 70-80% of the total. The molar concn of the blood of the snail steadily

increases ($\Delta > -0.1^\circ$) until rain or dew allows an imbibition of water. If this hydration does not occur, the animal dies when Δ reaches about -0.6° . The total blood CO_2 varies inversely with NaCl concn during the state of activity and constitutes 20–30% of the total molar concn. NaCl and carbonates account for almost the entire molarity of snail blood. It contains only 2.3 mg of mineral P per l. H. EAGLE.

Goldfish susceptible to hydrocyanic acid gas. R. J. RYAN, Calif. Dept. Agr., *Monthly Bull.* 10, 728 (1930). In 2 citrus orchard fumigation with liquid HCN caused the death of goldfish colonies in ponds located about 100 ft. from the nearest tree. Cf. C. A. 14, 2051. C. R. FILLERS.

Metabolism of the honey-bee colony during winter. C. I. COCKING, Wyoming Agr. Expt. Sta. *Bull.* No. 175, 3–51 (1930). As the external temps. decline below the clustering temp., the metabolic rate is not materially increased. C. R. FILLERS.

Cholesterol of decapod crustaceans. A. FRIEDMAN AND A. POLICARD, *Compt. rend. soc. biol.* 103, 823 (1930). *Physiol. Abstracts* 15, 112. A method of extg. the cholesterol from certain crustaceans is described. The cholesterol is similar to that of higher animals but its rotatory power is somewhat higher. G. G.

Influence of thyroxine on growth of embryos of the pond snail. O. V. IVKINA, *Compt. rend. soc. biol.* 103, 6118 (1930). *Physiol. Abstracts* 15, 180. Thyroxine has a depressing influence on the growth of the embryos of the pond snail (*Physa*). It alters the normal process of differentiation but has no direct influence on the duration of development. G. G.

Ion action. IV. The influence of various ions on the rate of the load heart. SHUNJI HONMA, *Japan J. Med. Sci.* 1, No. 3, 101–17 (1929). Expts. were conducted on the isolated toad heart perfused with oxygenated Ringer soln. The NaCl in it was replaced by salts contg. NaBr, NaI, etc. in various concns. It was found that rubber tubing must not be used in the work. Br and I ions show a positive chronotropic action. In weaker concns this is not evident. Weak concns of NO_3 ions show a positive chronotropic effect, but strong concns are markedly negatively chronotropic. SCN ions are practically the same as NO_3 ions. SO_4 ions are very weak and chiefly negative. Anions in a given concn are very different and the sensitivity of the heart beat for these anions shows a lyotropic series. Cl shows a negative chronotropic action, and in concn of 0.11 mole the heart is quickly stopped. An antagonism exists between Li and K. V. The influence of various ions on the excitability of heart strips of the frog. *Ibid.* 147–55. Expts. on strips of heart muscle from *Rana nigromaculata* in a Ringer strip preparation in which various ions were interchanged showed certain changes in irritability when stimulated with direct current of const. duration. Br, I and SCN ions in weak concns increase the irritability. In high concns the irritability is more or less decreased. SCN shows the greatest effect. SO_4 produces a diminution, while NO_3 ions cause an increase. The concn. of each anion for max. increase in irritability also shows a lyotropic series. The influence of anions on contractility is independent of the irritability. G. H. W. LUCAS.

Intracellular oxidation-reduction studies. III. Permeability of echinoderm ova to indicators. ROBERT CHAMBERS, BARNETT COHEN and HENRIET POLLACK, *J. Exptl. Biol.* 7, 1–8 (1931).—Ova of the combellar (*Echinarachmus purma*) and starfish (*Asterias rubens*) were fertilized and freed from jelly and membranes before they were immersed in solns. of indicators. Eggs that were unaltered were injected with 1% K₂Cr(CN)₆ as a test for reduced, colorless indicator. Toxicity of the various stains was gaged by the subsequent development of the eggs. All indophenols studied penetrated the egg membranes except those with a sulfonate radical. Under aerobic conditions the indicators of the series stained down through methylene blue are reduced within the ova. Penetrating indicators are toxic except phenol blue sulfonate; the indicators that failed to penetrate were non-toxic. Upon cytolysis or tearing, the indicator is washed from the ovum and the reducing substances pass into the surrounding fluid but retain their activity. C. M. McCAY.

The effect of hydrogen-ion concentration on protozoa, as demonstrated by the rate of food vacuole formation in *Colpidium*. SYLVIA M. MILLER, *J. Exptl. Biol.* 7, 17–20 (1931).—*Colpidium* cultures (from hay infusion) were suspended in a series of buffer solns. and fed under mk. suspensions. The rate of formation of food vacuoles indicates the feeding rate. This attains a max. at pH 6.0. The max. rate of movement also occurs at this pH. The mucus produced for food collection has a max. viscosity at a pH of 8.0. This value lies in the zone of decreased food consumption and movement. C. M. McCAY.

The adaptation of *Gundia ulvae* to salinity. I. The environment. C. F. A. PANTIN, *J. Exptl. Biol.* 7, 63–72 (1931).—*Gundia ulvae* occurs at the mouths of streams

flowing into the salt water of the ocean. It must withstand changes from fresh water to undiluted sea water. It may be exposed to either for periods of several hrs. II. The water exchange. *Ibid* 73-81—These worms immersed in dil. sea water swell. At great dilns. the swelling is less than if they were covered with semi-permeable membranes. In waters contg. only small amts. of inorg. elements the worms swell and die. If these fresh waters are rich in CaCO_3 , the swelling is much less. NaCl , NaHCO_3 , or glycerol cannot be substituted for the CaCO_3 . III. The electrolyte exchange. *Ibid* 82-94—The loss of salts upon the transfer of the worms from sea water to fresh water was measured by elec. cond. Twenty-five % of the salts are lost as the worms imbibe water to double their vol. In waters contg. CaCO_3 , in which the worms survive their internal salt concn. is reduced to 6-10% of that of sea water. Ca alone seems able to check this salt loss since it cannot be replaced by the other electrolytes studied and since salt loss is continuous in distd. water up to the stage of cytology. C. M. M.

Researches on the interior medium of the marlin lamprey (*Petromyzon marinus*)—its variations as the exterior medium is changed. FONTAINE *Compt. rend.* 191, 680-2 (1930)—Lampreys were placed in fresh water which was progressively dild. with sea water. After 24 hours' immersion in water of a given concn. the Cl content and osmotic pressure of the blood were detd. The Cl concn. increased but the ratio of Cl to other ions in the blood decreased. Mucous retards the appearance of pathol. symptoms when the lamprey is placed in salt water. J. G. McNALLY.

A comparison of the action of magnesium chloride on the respiratory activity and on the contractility of the frog ventricle. J. LEIDOWITZ AND A. SCHWITZER *Biochem. Z.* 229, 291-5 (1930)—At the close of the seasonal hypersensitivity of the frog heart to MgCl_2 it is found that Mg salts exert an inhibitory effect upon tissue respiration which is parallel to the diminishing inhibition of the contractility. The inhibiting action of small doses of Mg on the oxidation can be shown during the summer in the heart but not in the liver. S. MOSCOWITZ.

Action at a distance and the development of the eggs of the sea urchin. New experiments. J. MACROU AND Mme M. MACROU *Compt. rend.* 191, 802-4 (1930), cf. *Ibid* 186, 802 (1921), 188, 733 (1922), 189, 770 (1923)—Eggs of the sea urchin (*Paracentrotus lividus*) are markedly affected at a distance by exposure to bacterial suspensions and chem. mixts., although sepd. therefrom by a quartz flask with a ground and hydraulically sealed stopper. These expts. were repeated with controls not exposed to the above agents, and slight irregularities of growth in the controls suggested that the larvae do not perfectly accommodate themselves to the sealed flask. A new set of expts. was therefore performed in which the active substances rather than the eggs were placed in the closed vessel. The results confirmed the action at a distance effected by *B. tumefaciens*, and glucose in an oxidizing medium. The open top of a tube contg. the active substance may be bent down and introduced into a control tube without effect on the development of the larvae, thus showing that the action at a distance is not due to a substance capable of traveling through this tube. T. H. RIDER.

Actions at a distance on fertilized eggs, sperm and virgin eggs of the sea urchin (*Paracentrotus lividus* Lk.). J. MACROU AND Mme M. MACROU *Compt. rend.* 191, 963-5 (1930), cf. preceding abstr.—Sperm of *Paracentrotus lividus* Lk. exposed in quartz vessels to the same conditions as reported in the previous article, when used to fertilize normal virgin eggs, gave larvae which developed abnormally. Virgin eggs likewise treated before fertilization gave abnormal larvae. The exposure of eggs or sperm in the same manner to a soln. of uranum nitrate too weak to affect a photographic plate in 24 hrs. still produced abnormalities. T. H. RIDER.

Chemical processes of cellular division. LOUIS RAPIKIN *Compt. rend.* 191, 871-4 (1930)—Accepting the hypothesis (HARRIS, *Proc. Roy. Soc. (London)* B94, 426 (1923)) that cellular division is preceded by a denaturation of proteins accompanying a splitting of the cellular lipid proteins, R. investigated the occurrence of the S11 radical (I) during the stages before and after the division of eggs of *Paracentrotus lividus* Lk. Virgin eggs showed 35 mg. I per 100 g. eggs. Thirty min. after fertilization this figure dropped to 10 mg. I, and 10-15 min. before the first division increased to 46 mg. I. Eggs exposed to the action of 0.00001 M HgCl_2 20 min. after fertilization for a period of 5 min. then transferred to sea water or sea water contg. PhNH_2 or cystine failed to develop. Transferred to a 0.001 M soln. of cysteine in sea water, however, 70-90% developed. Thus cellular division can be controlled by the removal or addn. of L. T. H. RIDER.

Variation in the response to light in *Ameba proteus* with special reference to the effects of salts and hydrogen-ion concentration. S. O. MAST AND H. R. HULFIRU *Protoplasma* 11, 412-31 (1930)—The rate of reaction is largely dependent upon age

size, etc., of the ameba. Under comparable conditions increase in HCl causes increase in reaction time. The reverse effect is obtained with CO_2 . An increase in concentration of the chlorides of K, Mg and Ca causes a decrease in reaction time. In solns. of NaCl there is no consistent relation. M. H. SOULE

12—FOODS

F. C. BLANCE AND H. A. LEPPER

Amendment to the Federal food and drugs act of June 30, 1906. Food and Drug Administration, U. S. Dept. Agr., *Service and Regulatory Announcements*, F. D. No. 1, Suppl. 1 (1930).—The text of the amendment of July 8, 1930, to Section 8 of the act is given. The Secretary of Agriculture is authorized to promulgate, from time to time, such standards of quality, condition and (or) fill of container for canned foods as will in his judgment promote honesty and fair dealing with the consumer. Canned food coming within the jurisdiction of the act (except meat and meat food products subject to the Federal Meat Inspection Act, and canned milk) which fall below such standards is misbranded unless its package or label bears a plain and conspicuous form of statement promulgated by the Secretary, indicating that it falls below the standard. Canned food is defined as "all food which is in hermetically sealed containers and is sterilized by heat." It is provided that nothing in the amendment shall be construed to authorize the manufacture, sale, shipment or transportation of adulterated or misbranded foods." H. A. LEPPER

Definitions and standards for food products. Food and Drug Administration, U. S. Dept. Agr., *Service and Regulatory Announcements*, F. D. No. 2, Suppl. 2, August, 1930.—Definitions and standards are adopted for fruit juice, grape juice, orange juice and whole wheat flour. The standard for flour has been revised. *Ibid* Suppl. 3, January, 1931.—In the definitions and standards for food products adopted as a guide for officials of the department in enforcing the Federal food and drugs act the term "sugar," with or without the parenthetical expression "sucrose," as used to designate the sweetening agent in manufactured food products, is to be interpreted to include dextrose (pure refined corn sugar). This is in keeping with the announcement of the Department permitting the use of pure refined corn sugar as an ingredient of food in which sugar (sucrose) is a recognized element without declaration on the label of the product. The restrictions on the adulteration or imitation of any natural product, such as honey, by the addition of any sugar are not affected by the ruling. H. A. LEPPER

Crude fiber determination with the aid of a filter of a new type. K. FEIST AND E. KUNTZ. *Z. Untersuch. Lebensm.* 59, 480-3 (1930).—An app. is described by which dependable results are obtained in min. time. C. R. FELLERS

Detection of hard wheat grits. E. BERLINER AND R. RÜTER. *Z. ges. Getreidew.* 16, 212-3 (1929).—A criticism of Kühn's test (*C. A.* 24, 4557). B. C. A.

Detection of hard wheat grits. H. KÜHN. *Z. ges. Getreidew.* 16, 213-6 (1929); cf. *C. A.* 24, 4557.—A reply to Berliner and Rüter (preceding abstract) and statement of exact conditions. B. C. A.

Effect of harvesting in different stages of ripening on the valuable constituents of cereals. F. BERNER AND W. SCHLIMM. *Landw. Jahrb.* 72, 269-61 (1930).—Late harvested grain has greater weight per 1000 grains. There is a migration of the starch to the grain in the later stages of ripening and a rapid increase in the protein content between the "milky" stage and ripeness. GEORGE R. GREENBANK

Biometric analysis of cereal-chemical data. I. Variation. ALAN E. TRELOAR. *Cereal Chemistry* 8, 69-88 (1931).—T. explains the fundamentals of statistical analysis. He gives several formulas which are used by the biometrician. Histograms are given with fitted curves superimposed. The significance of statistical consts. is explained. L. H. BAILEY

Baking value of flour and possibility of determining it in the laboratory. A. TASMAN. *Chem. Weekblad* 27, 138-43 (1930).—An account is given of the attempts to correlate chem. and phys. consts., particularly viscosity and swelling power, with baking properties. Valuable information may be obtained by detns. of viscosity in the presence of regularly increasing quantities of acid, the differences in viscosity being plotted to form a curve. From the form of the curve for a given flour, rough estimates may be made of the ash content, gluten content and swelling power. B. C. A.

The stability of leavening in self-rising flour. PAUL LOGUE AND ELIZABETH MCKIM. *Cereal Chemistry* 8, 24-9 (1931).—The consumption of self rising flour in

the U. S. is increasing more rapidly than the consumption of plain flour. The rate of premature decomposition of self rising flour varies directly as the moisture content. Self rising flours containing less than 12.5% moisture are relatively stable. Self rising flours containing more than 12.5% moisture decompose more rapidly and result in alk. baked products, unless this premature decomposition is compensated for by increased acid leavening agent. An explanation of the increased alk.-accompanying premature decomposition is offered in the mechanism of the reactions involved. L. H. BAILEY

Relation of peptization of wheat flour protein to loaf volume. R. H. HARRIS *Cereal Chemistry* 8, 47-63 (1931).—A series of 11 experimentally milled flours was peptized by 3 methods and the percentage of total protein extract was computed. The correlations between improver loaf vol and percentage of total protein peptized were also computed. The quantity of protein removed increased quite rapidly at first, but no appreciable change was detected after the initial 30 min. Temp did not appear to have any significant effect upon the quantity of protein extract except in the water suspension, when a slight increase with rise in temp was noticed. Forty four flours of 75% extn., milled from 1929 crop Saskatchewan Marquis wheat of contract grade, were peptized by 0.5 N soln. of KBr and $MgSO_4$, and the protein extract was rated as percentage of the total protein, also as peptized and non peptized protein fractions of the flour. Correlations between the percentage of total protein peptized and improver loaf vol were lower than those between total protein and improver loaf vol and show less probability of forecasting loaf vol. The relationship between non peptized protein and improver loaf vol is not significantly different from that involving total protein and improver loaf vol, and is of equal value in predicting baking strength. The corresponding correlations obtained with the basic loaf vol were not so large, and are relatively of less practical importance. Twelve flours milled in a similar manner from 1923 pure strain Marquis grown at various Saskatchewan points including grades from No. 1 Northern to No. 6 gave a higher neg. correlation between the percentage of total protein peptized and improver loaf vol than between the percentage of total protein peptized and basic loaf vol. L. H. BAILEY

Predetermining flour soundness. D. A. COLEMAN *American Miller* 58, 434 (1930).—In an endeavor to find some method of detg. the soundness of flour without going through the time-consuming methods of milling and baking, it was observed that the acidity of the fat is much greater in unsound wheat than in normal wheat. It was also observed that the normal acidity of the different classes of wheats varies widely. Thus the acidity of the fat of sound hard red winter wheat does not exceed 8.5, hard red spring wheat, 5.5, soft red winter wheat, 8.0, Durum wheat, 4.0, and white wheat, 5.5. The acidity of the fat of certain types of damaged wheat was found to be dank and wet wheat, 8.6-16, musty wheat 9-15, sour wheat, 9-22, skin burned wheat 21-34 and mahogany, bin burned wheat, 35-44. The method of testing the acidity consists in removing the fat from the material with petroleum ether, recovering the fat from the ether to det. the quantity present and dissolving the fat in a suitable solvent so that the free fatty acid can be detd. The acidity of the free fatty acids present is expressed per g. of fat. L. H. BAILEY

Wheat and flour studies. XVII. Factors influencing the viscosity of flour-water suspensions. 4. The effect of extraction with solutions of potassium and sodium halides. ARNOLD H. JOHNSON *Cereal Chemistry* 8, 39-43 (1931), cf. *C. A.* 24, 3287.—The viscosity of flour water suspensions extract with 1.0 and 0.5 N solns. of the K and Na halides are higher than those of similar suspensions extract with dist. water. The viscosities decrease according to the lyotropic series of the anions $I > Br > Cl > F$. Glutenin is not removed on extract of a flour suspension with the halides of K and Na. L. H. BAILEY

The march of acidity in stored flours. II. M. C. MARKLEY AND C. H. BAILEY *Cereal Chemistry* 8, 29-38 (1931) cf. *C. A.* 24, 897.—A higher correlation of ash content with acidity as detd. by the A. O. A. C. method was found in freshly milled flours than the ash content with acidity detd. by the Greek method. The difference between Greek acidity and A. O. A. C. acidity in freshly milled flours was highly correlated with the ash content. Earlier assumptions that the A. O. A. C. tentative method for acidity indicated the relative rate of production of acid reacting substances during the extract with water received support from these recent studies. The rate of change in acidity of flours stored at 25° was highly correlated with ash content. It was extract that flours containing the percentages of ash included in the 4 grades or qualities described in the present Greek specifications will contain in excess of the max. limits of acidity after 43 to 52 days when stored at 25°. L. H. BAILEY

Recognition and determination of the polyfructoses. II. Determination of tri-

fructose anhydride for estimating the rye content of flour. C I KRUISHEER. *Rec. trav. chim.* 50, 153-63 (1931). cf C A 24, 4572—Treat 12.5 g of flour with 50 cc. of 70% alc and shake vigorously for an hr in a closed flask. Filter and transfer 2% cc. of filtrate to a centrifuge tube and treat with 5 cc of N NaOH in 70% alc. Mix and allow to stand overnight. Centrifuge, pour off the liquid and wash the residue twice with 2 cc portions of 70% alc. Dissolve the ppt in 10 cc of hot water and neutralize to methyl orange with 0.25 N H_2SO_4 . Transfer the soln together with 15 cc more of water to a 50 cc measuring flask, add 2.5 cc of 9.5 N HCl and heat 10 min at 68-70° to invert the sugar. After cooling, neutralizing and again cooling, the soln is mixed and in 10 cc of the soln, the total reducing power (R_1) is detd preferably by the method of Schoorl. To det the fructose content (F_2) take 20 cc of the inverted liquid and treat with hypiodite. From the values R_1 and F_2 , the content in trisfructose anhydride (T) can be computed and also the saccharose obtained. Take the liquid poured off after the centrifugation described above, make neutral to methyl orange and evap. to a small vol. Transfer to a 50 cc measuring flask, dil. to about 2% cc., add 2.5 cc of 9.5 N HCl and invert 10 min at 68-70°. Continue as described above to det R_1 and F_2 . The results in the analysis of about 40 samples are shown. In bolted rye meal 1.5-2% of a trisfructose anhydride was obtained, but only 0.1-0.3% in wheat flour. On standing the content increases. On the basis of the above behavior, the rye content of flour can be estd to within about 10%. The necessary calcs are described. W T H

Determination of butter fat. E SCHWEITZER AND J GROSSFELD. *Z. Untersuch. Lebensm.* 59, 491-501 (1930).—For 3 yrs a large no. of samples of butter, mixts of butter and margarine, and margarine were examd for butyric acid and Reichert Meissl values. The mean % deviations from the normal values of 20.0 and 27.0 were ± 7.4 and ± 10.0 , resp., for normal butter. The mean of the Reichert Meissl values for 37 samples was 24.7, which corresponds closely with the van Raalte value of 24.4. The ratio of the butyric acid value to the Reichert Meissl value was in excess of 0.60 in all cases (av. 0.713). Other relations are discussed. C R FELLERS

Masking of the color reaction of margarine. B HEPNER AND S ZALC. *Przemysl Chem.* 14, 412-5 (1930).—Under some conditions the use of oil of sesame for distinguishing between butter and margarine is not satisfactory because of the difficulties encountered in the Baudouin reaction. It depends on the presence of sesamol in the oil of sesame which may be destroyed during refining either by acid reagents or high temp (180°). Rancidity of the oil of sesame or of margarine produces under the influence of HCl an intensive coloration which masks the Baudouin reaction. This reaction was found unreliable in all cases of rancid margarine. KOH should be used for the removal of phenolic dyes used for coloring margarine when these mask the furfural reaction. A C. ZACHLIN

Determination of the milk proteins. II. GEORGE M MOIR. *Analyst* 56, 73-8 (1931), cf C A 25, 1004.—On the basis of the isoelec. point it is suggested that casein be defined as the material pptd from cow milk at pH 4.6 by AcOH buffered with NaOAc. By digesting with N NaOH and detg the increase of the formal titration and by studying the rate of oxidation, it is shown that the casein pptd under the above conditions is identical with that pptd by AcOH alone at pH 4.2. W T H

Number and character of bacteria and their subsequent development in pure and in contaminated milks. HENRI L DEFRARD. *Sci. Agr.* 11, 369-77 (1931).—Contamination of milk with manure or soil bacteria not only augments the initial nos. of organisms present, but these types develop in the milk during several days storage. Temp. increases during storage of contaminated milks cause great increases in soil and manure bacteria in milk. A contamination of approx. 0.1 g of soil or manure per l does not produce marked changes in the character of the milk flora if the temp. is held at about 60°F. If held below 50°F there is a considerable increase in the no. of liquefiers and alkali producers; this is due principally to staphylococci. C R FELLERS

Modification of the methylene blue reduction test and its comparative value in estimating keeping quality of milk. C K JOHNS. *Sci. Agr.* 11, 171-90 (1930).—The suggested modifications to the standard test are: preliminary incubation at 12.8° for 18 hrs. and mixing of contents of tubes not decolorized in 6 hrs. when subsequently incubated at 37.5°. These modifications shorten the reduction time, and the mixing reduces the variations between duplicate tubes. The chief advantages are greater convenience, improved accuracy on high grade milks, and closer correlation with keeping quality. In a study of 145 samples by the ordinary and the modified reduction tests, acidity increase, and plate and breed counts, the modified reduction test proved to be the method best suited to aid in ranking a series of samples according to keeping quality. C R FELLERS

Keeping quality of combined fat extraction, fixing and staining solution used in the direct microscopic technic for counting bacteria in milk. R. W. NEWMAN, Calif. Dept. Agr., *Monthly Bull.* 19, 607(1930)—No deteriorative changes were apparent after 3 yrs. use. The bottles were kept well stoppered. C. R. FELLERS

Investigation of the occurrence of *B. abortus* (Bang) in the milk of English herds (county of Somerset). D. R. WOOD AND E. T. ILLING. *Analyst* 56, 105(1931)—The deposit obtained by centrifuging 100 cc. of milk was injected subcutaneously into a guinea pig. After 4 weeks the pig was killed and cultures were made from the interior of the spleen on glucose, neutral red and vitamin agar in an atm. of 10% CO₂ and 90% air at 37°. The glucose and neutral red show at once whether any growth obtained is due to contamination with *B. coli*. *B. abortus* was identified by its failure to grow except under reduced O₂ tension and by agglutination with standard anti serum. The expts. were carried out concurrently with the examn. of the milk for tubercle bacilli (C. A. 24, 5387). *B. abortus* was found in the milk of 4 out of 379 herds tested. W. T. H.

Effect of the electropure process of treating milk upon bacterial endospores. A. J. GELT, JR. *Proc. Assoc. Southern Agr. Workers, 31st Ann. Convention* 162-7 (1930)—Although the cultures originally showed 80 to 90% sporulation no spore-formers were found after the samples were subjected to the electropure process. Further expts. indicated that purification of milk by the electropure process is brought about by the aid of another factor besides heat. K. D. JACOB

Retained milk. OTAKAS LAKA. *Ann. feldt* 23, 609-10(1930)—Four cows which had been unmilked for 5 days furnished on milking only about 100 cc. of thick, brownish tinged, colostrum like milk, having an odor of burnt milk and the following compn. water 79.40, fat 10.15, casein 2.65, albumin and globulin 0, albumoses and peptones 2.03, amino acids 1.78, lactose 1.50, ash 1.10, undetd. 1.50%. The ash contained 23.2% CaO and 16.76% Cl. A. PATINEAU COURTES

Investigation of artificial whole-milk powder. P. WEINSTEIN. *Z. Untersuch. Lebensm.* 59, 515-6(1930)—Microscopic examn. of a powder showed a relatively high proportion of large fat globules. Extn. of the fat with light petroleum ether gave a yield of 19.0%. The n_D and other consts. showed the oil to be largely olive oil which had been used for adulterating skim milk powder. C. R. FELLERS

Studies concerning the daily yield and fat content of the milk of the half-wild yak in the Siberian Altai Mountains, together with a contribution to the investigation of milk and fat inheritance in primitive breeds. J. W. ANSCHUTZ. *Z. Tierzucht. Zuchtungsbiol. Tierzucht* 20, 293-308(1931)—Studies are reported of the yield and fat percentage of the milk of 116 yaks and 6 yak crosses with Altai cattle living on sparse pasture in Siberia. The daily milk yield of the yak cows varied from 1.2 kg. to 4.1 kg. with a mean of 2.73 kg. The fat varied from 4.23 to 8.50% with a mean of 6.07%. The range of yield for the yak crosses was 2.4 to 4 kg. with a mean of 3.20. The fat varied from 4.59 to 6.21% with a mean of 5.45%. L. A. MAYNARD

Non-acid Babcock method for determining fat in ice cream. O. R. OVERMAN AND O. F. GARRETT. *Ill. Agr. Expt. Sta., Bull.* No. 360, 393-406(1930)—The method requires 2 reagents, A and B. A is made by mixing 75 cc. of c. r. NH₄OH, 35 cc. of BuOH and 15 cc. of 95% EtOH (denatured is satisfactory). Reagent B is made by dissolving with heat 200 g. of Na₂PO₄ and 150 g. of NaOAc in 1 l. of water. To make the test, weigh 9 g. of the well mixed sample into a 8 or 10% milk test bottle (or a 20% ice cream test bottle), add exactly 2.5 re. of A and mix thoroughly. Then add 9-10 cc. of B and thoroughly mix. Place the test bottle in a shallow water bath and heat to boiling, continuing the heating for several min. During the heating shake the bottles several times. The fat seps after 15-30 min. and forms a clear yellow layer at the top of the liquid. After the fat sepn. is complete, place the test bottle in the centrifuge and whirl successively, as in the Babcock test, for 5, 2 and 1 min. Add hot water, as in the Babcock test, but do not acidify the water. Then place the bottle in the hot water bath (130-40°F) for 5 min. In reading the results measure from the bottom of the fat column to the top of the upper meniscus. Comparisons were made with the Official Rose-Gottlieb method. Of the 54 samples tested, 64.82% came within ±0.1% of the R. G. values 24.07% fell within ±0.1-0.2 and 11.11% between ±0.2 and ±0.3%. The av. difference for all samples was ±0.0675%. The method is fairly rapid, inexpensive and sufficiently accurate for com. use. A large no. of detns. can be made concurrently. Twenty-six references are cited. C. R. FELLERS

Standardization of milk with skim milk for the manufacture of Cheddar cheese. H. C. HANSEN AND D. R. THORPHEUS. *Idaho Agr. Expt. Sta., Bull.* 174, 3-16(1930)—See C. A. 25, 360. C. R. FELLERS

The detection of horse meat as an adulterant in sausage and other studies of the precipitin test. F. M. BOLIN *J. Am. Vet. Med. Assoc.* 31, 103-70(1931).—A horse meat ext. antiserum was found better than horse serum antiserum for the detection of horse meat adulteration in sausage. Rabbits immunized to a cooked product had sp. precipitins for the uncooked sp. protein, and at the same time had nonsp. precipitins for cooked meats of different species. Meats tested at 15 lb. pressure for 10 min. were unsuitable as antigens for immunization purposes. Immature animals were unsatisfactory in the production of precipitins. Meats of closely related species were differentiated only by the length of time necessary for the reaction to occur.

RACIEL BROWN

"Adulterated" olive oil in sardines. H. WINKER *Z. angew. Chem.* 43, 626 (1930). High I values of the oil used in prepg. sardines, thought to be due to adulteration of the olive oil, were found to be associated with the extn. from the sardines of a fat of very high I value by the pure oil.

R. C. A.

Increase in the p_{H} of the white and yolk of hen eggs. PAUL FRANCIS SHARP AND CHARLES KELLY POWELL *Ind. Eng. Chem.* 23, 1960(1931). cf. *C. I.* 24, 4338.—The p_{H} of eggs is an important factor in detg. their rate of deterioration. Eggs with a low p_{H} deteriorate less rapidly. The p_{H} of the white increases rapidly as the result of an escape of CO_2 gas. This may be delayed by storing at low temp. or by oil dipping. After 15-25 days storage at high temp. the p_{H} decreases. The p_{H} changes of the yolk are smaller. In com. incubators the p_{H} of the white after 2 days is 8.02, which may affect the developing embryo. A higher CO_2 pressure would remedy this.

WYLLIE LEVINSKY

Preservation of fruits and vegetables by freezing storage. M. A. JOSELYN *Calif. Agr. Expt. Sta. Cir.* 320, 33 pp (1930). cf. *C. I.* 23, 3630.—At temps. of $10^{\circ}F$, molding, fermentation and bacterial spoilage do not occur and the normal ripening and respiration processes are effectively arrested. The com. methods are outlined. Deteriorative changes due to oxidative enzymes go on slowly at freezing storage temps. and cause surface discoloration in some fruits. The usual ratio of fruit to sugar is 2 or 3 to 1. A syrup pack is preferred to a mist of fruit and dry sugar, as surface discoloration is thus greatly reduced and better distribution of the sugar is secured. After thawing, cold packed fruits keep from 3 to 7 days without actual spoilage. Vacuum packing of fruit for freezing possesses the advantage that nearly all the O_2 is excluded from the container, thus decreasing oxidative changes in the frozen product. Fruit juices may be frozen without much injury to their flavor. Though most vegetables may be successfully stored in the frozen state, slight changes in flavor and the inability to keep for more than a few hours after thawing render frozen vegetables less suitable for com. exploitation than fruit.

C. R. FILLERS

Storing canned fruits at high temperatures injures quality. R. Y. KOSMAN *Food Ind.* 3, 77(1931). cf. *C. I.* 24, 1679.—The hydrolysis of the anthocyanin pigments in red fruits is greatly increased at high storage temps. The color base separates from the glucoside base and is thrown out of soln. as a brown amorphous mass which no longer retains any of the original color. This mass causes some darkening of the fruit, but more important is the loss of bright natural color. Prompt and efficient cooling after the heat treatment, as well as subsequent cool storage, is strongly recommended as a means of better preserving the color of fruits.

C. R. FILLERS

Research answers (food) packing and bottling problems. I. C. G. FICKS *Glass Container* 10, No. 4, 8-9, 22-30(1931).—Animal feeding tests for vitamin C in 1930-31, sweet prep. in various ways showed that for max. retention of vitamin C, complete air removal is necessary. Tomato juice packed without treatment to remove air was entirely devoid of vitamin C. Vacuum sealing of the container was ineffective, but vacuum treatment of the juice before bottling combined with vacuum sealing, proved most effective in conserving the vitamin C in the final product. By decreasing the particle size by homogenization, sepn. of the pulp and juice was largely prevented. Light-oz. bottles of tomato juice were effectively sterilized when subjected to a holding period of 15-20 min. in a water bath.

C. R. FILLERS

Disinfecting power of bleaching powder with special reference to its application to the disinfection of vegetables. KENJI BANNAI AND HIDEKICHI HABU, *J. Pub. Health Assoc. Japan* 6, No. 7, 3-6(1930). *U. S. Pub. Health Eng. Abstracts* 11, 11, 2(Jan. 31, 1931).—The disinfection with bleaching powder of such vegetables as greens, radishes, turnips, spinach, etc., could not be satisfactorily accomplished, irrespective of the concn. of the soln., length of exposure, no. of bacteria attacked, or time elapsed after contamination of the vegetables.

C. R. FILLERS

Storage of potatoes. N. D. VYAS *Agr. J. India* 25, 408-16(1930).—Loss of

potatoes during storage for 32 weeks was reduced to 28-31% by wt. by storage in cinders or charcoal as compared with a loss of 40-50% by storage in coarse sand. Addn. of lime slightly lowered the preservative power of the storage media. The preservative action of cinders and charcoal seems to be due to their ability to absorb moisture and gases given off during rotting of the potatoes, thus preventing infection of sound potatoes. Rotting of potatoes was increased by treating them with 5% H_2SO_4 for 3 hrs before storage. K. D. JACOB

Chemical studies on cannery peas. ANON. N. Y. Agr. Expt. Sta., 45th Annual Report, 50-1 (1930).—By means of a puncture tester the toughness of peas was found to increase regularly with size. Varieties varied considerably in toughness but the toughness of the fresh pea coincided with that of the processed (canned) pea as judged by experts. Chem. analyses were made for reducing sugars, sucrose, starch, sol. and insol. N, ash, Ca and alc. insol. residue on a large no. of samples. The following tentative phys. and chem. criteria are established for high-quality canning peas. The puncture value should not exceed 40 g. per sq. mm. The starch:sugar ratio should not be over 1.0. The insol.:sol. N ratio should not exceed 3.0 and the CaO content of the fresh skins should not be over 0.06%. Sucrase, diastase and protease are very active in macerated peas and may cause some deterioration in quality. Toughness and Ca content are correlated. Effects of fertilizer treatment on quality are being studied. C. R. FELLERS

Microscopy and histochemistry of certain fruits. ANNELESE NIETHAMMER, Z. Untersuch. Lebenst. 59, 501-6 (1930), cf. C. A. 24, 4076, 5890.—The endocarp of oranges contains crystals of CaC_2O_4 , which can be isolated and identified by treatment of the fruit with concd. H_3PO_4 followed by sublimation under reduced pressure. After storage for several weeks, slow resorption of the crystals occurs with concomitant formation of AcH . Lemons and mandarins exhibit identical phenomena. Pears contain characteristic inclusions, and in the ripe condition AcH , which is present in greatest abundance in over-ripe fruit. Upon sublimation, ripe pears give malic and citric acids, while over-ripe pears contain $H_2C_2O_4$ in addn. The AcH test is always pos. in plums, and CaC_2O_4 can be detected. From bananas in the ripening stage, malic acid can be sublimed, but after prolonged storage only traces of $H_2C_2O_4$ and tartaric acid can be detected. Unripe tomatoes contain considerable amts. of malic acid with traces of tartaric acid and $H_2C_2O_4$, while ripe samples contain malic, citric and tartaric acids. AcH is present only in ripe tomatoes. C. R. FELLERS

Papaya culture in Hawaii. W. T. POVE, Hawaii Agr. Expt. Sta., Bull. 61, 40 pp. (1930).—The chem. compn. of the ripe and immature fruits was detd. Improved varieties gave the following av. results: total solids 13.39, ash 0.76, acid (as H_2SO_4) 0.132, protein 0.484, sucrose 10.14, total sugars 10.61, fat 0.105 and fiber 0.902. The green fruit contains practically no starch or sucrose and only a small amt. of reducing sugars. The latter increase rapidly with approaching maturity. Papain, a powerful digestive enzyme, is present in the milky juice. Meat rubbed with a slice of green juicy papaya becomes tender as a result of the action of the papain. Because of the size, wt. and delicate structure of the fruit great care is necessary in harvesting, packing and shipping. A bibliography is given. C. R. FELLERS

Isolation of quinic acid from fruits. E. F. KOIDMAN AND N. H. SANBORN, Ind. Eng. Chem. 23, 126 (1931).—The BzOiff content of prunes and cranberries is too low to explain their acid reaction after digestion. Quinic acid, which the authors have isolated from these fruits and have detected in grapes, is more probably the cause. AMY LEVESCONIE

Sampling of apples for arsenical spray residue determinations. J. R. NELLER, Ind. Eng. Chem., Anal. Ed. 2, 382-4 (1930).—A statistical study of the sampling of apples that had been washed with dil. HCl showed that the av. probable error of single samples of 6 apples each was 8.2%, while that of duplicate samples was 5.36% of the total arsenic on the fruit. The av. probable error of analysis was 7.4%. The error of sampling is considerably reduced by using duplicate rather than single samples but the rate of reduction is much less for triplicates, quadruplicates, etc. J. R. NELLER

Artificial removal of astringency in kaki. KINZIRO KAKESITA, Proc. Imp. Acad. (Japan) 6, 397-8 (1930).—The expressed juice of Japanese astringent persimmons (kakishibu) is coagulated at room temp. by the addn. of AcH ($HCHO$ or $ClCHO$ may also be used). It is believed that this action is caused by the polymerization of tannic substances in the fruits with the formation of insol. tannin-colloids and that the astringency in kaki is thereby removed. C. J. WEST

Dehydration of grapes. P. F. NICHOLS AND A. W. CHRISTIE, Calif. Agr. Expt. Sta., Bull. 500, 3-31 (1930).—A new type of raisin grape called the 'Golden Bleached

Thompson" is the principal variety which is dehydrated, though wine and juice grapes can also be advantageously dehydrated. The grapes are dipped before dehydrating in a lye bath (0.5-1.0% NaOH) kept at 200-127°. In sulfuring the dehydrated product 2-4 lb. of S must be burned for each ton of fruit, giving a SO_2 content in the raisins of 200-1000 p.p.m. The foudling temp. should not exceed 16°-70° F. with a relative humidity of 25% or burning will result. Dehydration offers improvement in quality and yield by making it safe to defer picking until full maturity is reached. Special dips contg. olive oil, glycerol and mineral oil offer possibilities for improvement of color and keeping quality. Sulfuring did not adversely affect insect infestation. C. R. I.

Investigations on grape juice manufacture. ANON. N. Y. Agr. Expt. Sta., 49th Annual Report 55 (1930). The colloidal material in Concord grape juice affects adversely the color and clarity of the juice and delays the settling out of tartaric crystals. By removal of the colloids, crystallization takes place in a much shorter time, thus allowing the juice to be bottled. The colloidal matter in grape juice contains about 8% protein and 7% pectin, the rest being polysaccharides of unknown nature.

C. R. MILLER

The preservation of fruit juices, sweet must, etc. A. MEHLITZ. *Konserven Int.*, 1930, 742-4, 756-8. Cf. C. A. 25, 134. M. describes further work on the preservation of fruit juices by centrifuging the juices through a De Laval separator, clarifying by digestion with an enzymic ext. of *Aspergillus oryzae*, filtering through an asbestos pre-filter and then through a Sauter K. germ filter. It was necessary to reduce the active acidity of the red currant juice from pH 2.72 to 3.60, and of white currants from pH 2.69 to 4.15 to obtain better enzyme action. In every instance the juice treated with *Aspergillus oryzae* filtered much more rapidly than the untreated juice. After 4 months standing the filtered juices were clear and of good color and taste. F. P. G.

Study of the factors affecting the jellation of fruit juices and pectin solutions. P. B. MYERS AND G. L. BAKER. Del. Agr. Expt. Sta., *Bull.* 167, Ann. Rept. of the Director 22-(1930). With lemon albedo as raw material 4 series of pectins were prepared, 2 series each at temps. of 80° and 100°. The yield, jelly grade, jelly units, ash, methoxyl and pH of the pectins in each series were determined. The yield of pectin increased with the temp. of the extn., other conditions remaining constant. This condition existed only when the pH of the extg. medium was at 1.50 or above. When below 1.50 the yield of pectin increased with the temp. of the extn. up to some temp. below 100° and then decreased as the temp. was increased to 100°. For instance, when the extn. was conducted at pH 0.8, the yield of pectin increased from 20.5% at 10° to 30.8% at 85°, and then decreased to 24% at 100°. At some pH ion comes between pH 1.5 and 0.8 and below, the hydrolysis of pectin to some ale-sol. compounds is faster than the hydrolysis of protopectin to pectin at temps. below 100°. The yield of pectin at any temp. could not be increased to the optimum yield obtained at a temp. above it by increasing the time of extn. The optimum jelly grade of pectins, extd. by heating 30 min. while varying the pH of the extg. medium, decreased as the temp. of the extn. increased. Since the yield of pectin increased with the temp. of the extn. while the jelly grade decreased with increase in temp., the product of the 2 (jelly units) tends to remain constant so that the great variation in jelly units with the temp. of the extn. that was previously expected did not occur.

C. R. MILLER

Determination of sorbitol in jams. C. F. MUFFLET. *Ann. fals.* 23, 602-5 (1930), cf. C. A. 24, 3280. Cherry juice was found to contain 0.200% sorbitol, red gooseberry and strawberry juices contained none. Home made cherry jam contained 0.125%, while home made gooseberry jam and strawberry jam contained none. Commercial "pure fruit" gooseberry jam and "pure fruit" strawberry jam contained 0.135% and 0.075% sorbitol, resp., indicating the presence of apple juice, which was confirmed by the presence of malic acid. Commercial cherry jam contained 0.025% sorbitol, indicating that there was but little cherry (really only cherry fragments in sugar syrup which was jellied by addn. of com. pectin).

A. L'APENAU COUTURE

Determination of so-called crude fiber (cellulose) in cacao. K. KÖRSCHNER AND A. HANAK. *Z. Untersuch. Lebensm.* 59, 481-91 (1930). The method previously devised (Körschner and Hoffer, *Tech. Chem. Papier-Zellstoffabrikation* 26, No. 8-9 (1929)) for the detn. of cellulose in wood has been modified by the substitution of AcOH for HCl in the mixt. used for decomposing the noncellulose elements. The sample of defatted cacao, approx. 0.3 g., is intimately mixed with 15 cc. of 80% AcOH and 1.5 cc. of HNO_3 (d. 1.4) in a flask fitted with a ground in air condenser. The mixt. is gently boiled for 15-25 min. and then filtered through a large-pore porcelain or glass filter previously moistened with AcOH. The residue is washed successively with 7-10 cc. of the hot mixt. of HNO_3 and AcOH, hot water, a few drops of HCl , 5-10 cc. of

H₂O, 1-2 cc. of the acid melt, and finally hot water, until all traces of AcOH have been removed, the flask and walls of the filter being washed at each stage. The residue is dried at 100-8° and weighed. By this method, which gives most consistent results, white subgraded cellulose was extrd. from cacao and the detn. required less than 1 hr.

C. R. FILLERS

Determination of sulfur dioxide in foods containing sulfur-bearing volatile oils. FRANK A. DiBILLO. *Am. J. Pharm.* 103, 43-4 (1931).—The following method was used for detg. the SO₂ content of prepd. horse-radish and prepd. mustard to which had been added 150 p. p. m. of SO₂. To 54 g. of the material in a 550-cc. flask add 200 cc. water, 5 cc. 19% H₂O₂, and about 0.5 g. of tannin. Drill slowly in a current of CO₂ until 100 cc. is received in 20 cc. of soln. of H₂O₂. Det. sulfites with BaCl₂.

W. G. GAESSLER

Mineral content of pastures research. Some notes on the work in 1929-32. B. C. ARLOX. *New Zealand J. Agr.* 41, 10-7, 40-4 (1930).—An extensive investigation of the mineral content of N. Z. pasture grasses is being made in order to det., if possible, whether there is any relation between mineral deficiencies and various animal diseases which occur in specific areas. "Dropsy" in sheep seems to be connected with Ca deficiency in the Marion District, and it seems to be definitely established that "lock sickness" in cattle and sheep is caused by a deficiency of Fe in the pasture grasses of certain poorer soil areas. Deficiencies of I and P have been established in several localities. Numerous analyses of N. Z. pasture grasses are tabulated. K. D. J.

Investigations into the intensive system of grassland management. I. The chemical composition of intensively treated pasture. A. W. GREENHILL. *J. Agr. Sci.* 20, 513-8 (1931).—The chem. compn. of the dry matter of herbage from pastures managed under the intensive system and representing 2-3 weeks' growth was closely comparable with that obtained by other workers from "pasture cuts" at 1- to 2-weekly intervals from unfertilized pastures. The dry matter is a good feeding stuff of narrow nutritive ratio and low fiber content. The nutritive ratio is somewhat wider in a dry than in a wet season, but the av. fiber content is not markedly different. During the grazing season, drought or the early summer flush period produces some reduction in protein content and an increase in fiber content, with a recovery following cessation of the drought or the flush period. These variations are a reflection of the productivity and habit of growth of the herbage. An abundant growth of young leafy herbage has a higher nutritive value, wt. for wt. of dry matter, than the more starchy growth of early summer or drought conditions. The protein content of the dry matter of herbage from pasture intensively fertilized (including N) was higher than that from untreated pastures. The higher milk producing power of the former pasture may be ascribed to improved compn., as well as to higher yield, of the herbage. P. R. DAWSON

Nutritive value of pasture. VI. The utilization by sheep of mineral-deficient herbage. H. E. WOODMAN AND P. E. EVANS. *J. Agr. Sci.* 22, 507-517 (1930), cf. *C. A.* 24, 5432.—A detailed study was made of the utilization by sheep of a sample of herbage deficient in total ash, particularly in respect to Ca and P₂O₅. Na was also low, but the percentages of K, Mg and Cl were similar to those of "normal" herbage. The results lead to the belief that the failure of stock to thrive on mineral-deficient pastures cannot be explained by assuming that the low mineral content is responsible for lack of palatability and a consequent depression of appetite. There is no evidence that the shortage of minerals causes the herbage to be digested any less efficiently than "normal" cultivated herbage of similar maturity and org. compn. Maintenance on pastures of subnormal mineral content is due directly to failure of the diet to supply the necessary inorg. materials for constructive purposes and for maintaining the normal balance of minerals in the blood and tissues. P. R. DAWSON

Photography as a help in the examination of cattle foods. W. L. FIELDS AND S. T. PARKINSON. *J. South-Eastern Agr. Coll., Wyo., Kent No. 27, 45-72 (1931)*, cf. *C. A.* 24, 3575.—Detailed directions are given for the microscopical identification of leguminous plants and oil seeds in cattle foods. E. D. JACOB

Determination of acids inilage by Wiegner's method. E. GREEN. *Twines, Jr.* 1, 65-6 (1922).—A diagram permitting direct reading of the acids of AcOH and butyric acid has been constructed. B. C. A.

Polarimetric determination of starch in potatoes and the relation between potato dry substance and starch content (von Soest). 22. Composition [for treating containers for food] (Brit. pat. 235,559) 23. Emulsions of fatty substances [artificial cream] (Fr. pat. 950,242) 13.

WHITT, J. *State Med* 37, 559-70(1929); cf *C A*, 24, 5842—Evidence is given to show that most of the difficulties in the understanding of occupational skin diseases are due to the overemphasis of allergy and lack of knowledge of the common skin diseases. A. L. RAWLINS

Domestic mechanical refrigeration in relation to public health. ARNOLD H. KECFL. *Illinois Med J.* 58, 421-7(1930). *U. S. Pub Health Eng Abstracts* 11, 111A, 1(Feb 7, 1931); cf *C A* 24, 4561—In Chicago in less than 2 yrs there occurred 62 cases of poisoning with 10 deaths from the refrigerant gas MeCl . The serious casualties were traced to the multiple type refrigerator contg 200 lb of gas. The smaller household cabinets contain but 5 lb of gas. Refrigerating systems using NH_3 or SO_2 furnish their own warning when gas is escaping, whereas MeCl , EtCl and similar gases are practically odorless and may cause unconsciousness. Suggestions for safe installations are made. C. R. FELLERS

The measurement and regulation of humidity in industrial processes. FR. LIENEWEG. *Siemens Z* 11, 26-32(1931)—A review of modern practice. C. G. F.

Temperature distribution in internally heated cylinders. ALBERT B. NEWMAN. *Ind Eng Chem* 23, 29-32(1931)—A presentation of math theory of heat flow through walls of internally heated cylinders. Theory is applied to dissipation of heat generated by an elec wire carrying a current and insulated by an external covering. M. C. R.

Gases used in warfare. TH. MOARTEL. *J. pharm Alsace Lorraine* 57, 332-9(1930)—A review of the effects and chemistry of these gases and modes of protection. S. WALDNOTT

A useful horizontal-tank chart. W. F. SCHARHORST. *Ind Eng Chem* 23, 314-5(1931)—A chart is included which gives the no. of gallons of liquid in any horizontal tank. C. G.

Woods for insulation. A. R. DUNTON. *Electrician* 106, 237-9(1931)—The varieties used, methods of impregnation and breakdown voltages of treated woods are discussed. Typical phys. properties of the varieties of hard wood employed in insulation are tabulated, as are typical values of moisture content and elec strength for various woods. The woods are kiln-dried to a moisture content not exceeding 6%. Tables also show some woods used for bachelized veneer ply construction with typical tensile strength values of products, compression strength test results on these, and a comparison of elec strengths of untreated and bachelized veneer ply. The reliability of performance of all treated woods depends on their ability to withstand the penetration of moisture and it is therefore likely that thoroughly bachelized woods, impregnated in the form of thin veneers and built up into rods or boards, will prove popular. W. H. BOYNTON

The catalytic oxidation of CO (FRAZER) 2.

Annuario per le industrie chimiche e farmaceutiche Anno XI e XII. 1928-1929. Compiled by the Commissione per le industrie chimiche appointed by the Ministero delle Corporazioni. 724 pp. Paper, L. 100. Reviewed in *Chem Trade J.* 88, 128(1931).

CHAPLET, A. Dictionnaire des produits chimiques commerciaux et de la droguerie industrielle. Paris: Dunod. 308 pp. Paper, M. 54, bound, P. 63. Reviewed in *Ann fals* 23, 616(1930).

IZARD, J. Agenda Dunod, 1931. Physique industrielle. 11th ed. Paris: Dunod. 532 pp. Bound, P. 20. Reviewed in *Chimie & industrie* 25, 267(1931).

JAVET, E. Agenda Dunod, 1931. Chimie. 50th ed. Paris: Dunod. 612 pp. Bound, P. 20. Reviewed in *Chimie & industrie* 25, 267(1931).

LAUNAY, L. DE LA technique industrielle. Paris: Dunod. 337 pp. Bound, P. 85. Reviewed in *Chimie & industrie* 25, 267(1931).

PAGNIELLO, A. Disinfezione contro gli aggressivi chimici. Chimica di guerra. Rome: Istituto Santa Famiglia. 8 pp. Reviewed in *Chimie & industrie* 25, 260(1931).

STOLTZENBERG, HUGO. Anleitung zur Herstellung von Ultragiften. Hamburg: Norwi-Druck. 70 pp. M. 20.

STOLTZENBERG-BERGUIS, M. Was jeder vom Gaskampf und den chemischen Kampfstoffen wissen sollte. Hamburg: Chemische Fabrik Dr. Hugo Stoltzenberg. 30 pp. M. 1.

Adsorption of gases. Soc. chim. DE LA GRANDE PAROISSE (AZOTE ET PRODUITS CHIM.) Fr. 693,278, June 17, 1929. In a process for sep'g less volatile constituents

of a gas by solid adsorbent material, the gas is passed through the adsorbent under pressure and then expanded and the expansion is used through a heat exchanger to cool the gases being treated.

Separating mixtures of gases or vapors. *Soc. de Recherches et d'Exploitations Petrolières*. Ger 515,206, June 8, 1929. Mixts. of gases or vapors are passed through active C or other adsorbent and the adsorbed constituents recovered by means of steam after preheating the charged adsorbent. The improvement consists in preheating the adsorbent consecutively, e. g., by locally heating the adsorption app. so as to effect a circulation of the gases therein. App. is shown.

Separating gases by liquefaction. *Ges. für Friedr. Fischbacher A.-G.*, F. PATZER and H. KANT. Brit 2,55,599, June 20, 1929. Various details of app. and procedure are disclosed, suitable for the sepn. of constituents of purified coke-oven gas. Cf. C. A. 24, 2515.

Separating gas mixtures by liquefaction. *Société l'Air Liquide* (Soc. ANONYME L'Air de l'Exploitation des Produits GAZEUX CLAIRON). Ger 513,682, Dec. 1, 1929. Appl. to 495,429. In the method of Ger. 495,129 (C. A. 24, 3282) the cold unliquefied compressed residue is further cooled before it is caused to flow counter-currently in the mixt. Cooling may be effected by the expanded residue. The app. is shown.

Purifying liquefied gases. *N. V. Philips Gloeilampenfabriek*. Fr 692,625, Mar. 24, 1930. Liquefied gases such as O are purified by bringing them in contact with an absorbent material such as charcoal or NaO gel.

Deoxygenating air, etc. *Siemens-Schuckertwerke A. G.* Ger 515,133, Dec. 23, 1929. A mixt. of a metal and an electrolyte is used such that the metal, under the action of O, is converted into an oxidizable salt. A mixt. of Cu and NH_4Cl is specified. Catalysts, e. g., kieselguhr or active C and diluents may be included in the mixt. The method is intended particularly for eliminating O above the oil level of oil filled electric transformers, switches, etc.

Separation of solutions from precipitates. *H. LAWARRÉ*. Belg 371,101, July 31, 1930. An immiscible liquid, inert toward the reagents used or liberated, and having a different d from that of the solns involved, is used for the sepn. of solns. and ppt's. e. g., CCl_4 is used for the sepn. of the solns. in the mnt. of NaNO_3 and NH_4Cl from NaCl .

The separation of chemical solutions and precipitates. *H. LAWARRÉ*. Belg 371,580, Aug. 31, 1930. Solns. are sepd. from ppt's by means of a liquid which is inert toward the reagents used or liberated, immiscible with them and of different d . A fatty substance which can form a soap with the basic salts pptd. during the reaction is added to the liquid.

Aerating liquids. *WERNIG SUEH*. Ger. 507,204, Jan. 6, 1929. The air is dehydrated through perforated tubes encased in rubber, the latter having fine perforations which open only under the pressure of the air supplied.

Testing viscosity of liquid materials. *SARANT T. ROQUES*. U. S. 1,700,918, Feb. 3, 1928. A buoyant body is released at the bottom of a column of the liquid under test, permitted to rise through the liquid and the time required for its rise is noted. App. is described.

Obtaining pulverulent products from fused solids. *CHIM. FAN VON HADEN A.-G.* (Fritz Krumholz Inventor). Ger 514,110, Mar. 2, 1928. The solid is heated and stirred in a pressure vessel with a liquid in which it is difficultly sol. to a temp. above the m. p. of the solid and above the b. p. of the liquid. The mixt. is then discharged into a second vessel in which it is rapidly cooled by contact with a further amt. of liquid kept in rapid motion. Examples describing the treatment of β -naphthylamine and benzanthrol are given.

Emulsions of fatty substances. *ROBERT ROBERTS*. Fr. 693,242, April 3, 1930. Artificial cream and other emulsions rich in fatty substances are made by using condensed milk or other natural concd. soln. of phosphoprotein as protective colloid.

Electric insulators. *LOWE CORRELL, LTD.*, and *METATORS A.-G.* Brit 335,504, Feb. 18, 1930. For preventing moisture condensation on insulators such as those of *elec. gas-cleaning app.*, the insulator is heated from a suitable source of heat by a reflector which may be revolved around the insulator suitably with transmission of the heat through ordinary or quartz glass.

14—WATER, SEWAGE AND SANITATION

EDWARD BARTOW

Liege water service (1830-1930). E. THIRIAUX *Tech sanit munic* 25, 370-5 (1930) —An historical account. W. C. R. FELLERS

Kankakee (Ill.) served by new plant. Geo Husem. *Water Works Eng* 83, 478 (1930) —The new electrically equipped plant has a capacity of 6 million gals daily while the present rate of consumption is 2 million gals daily. River water is softened, clarified, carbonized, allowed to settle, filtered and made bacteriologically pure. Soda ash and lime reduce the hardness from an av. of 20 grains to about 4 or 5 grains per gal. The av. cost of chemicals will be about \$2000 a month. C. H. BADGER

How North Carolina develops its water works plant operators. H. E. MULLER *Water Works Eng* 83, 423-4, 446-9 (1930) —The different acts and laws passed in N. C. pertaining to sanitation and water supplies are reviewed. Inexperienced but technically trained chemists and engineers are secured through the Bureau of Engineering for duty at small plants at a salary of \$125 a month. Under the guidance of trained engineers these men are given thorough training in plant operation, serving a sort of internship for 3 yrs. until they can assume responsibility. In 1921 none of the 18 surface gravity supplies was chlorinated but in 1929, 19 out of 20 were. All filtered water supplies are now chlorinated. In 1921, 50 filter plants were deficient, 36 being obsolete, 10 requiring complete overhauling and 4 needing minor repairs. In 1929 all except 2 of the 78 filter plants were fully meeting modern practice. Only 1 plant was under technical supervision in 1921 while in 1929, 57 plants had complete chem. and bacteriol. control at the plant and 17 had simple lab. control of dosage. C. H. BADGER

A limnological significance of free carbonic acid. E. LINDEMANN *Naturwissenschaften* 18, 1113 (1930) —Evidence is advanced for the importance of the CO_2 content of water for the growth of *Pseudomonas* provided the pH is favorable. B. J. C. V. D. H.

Deciding type and layout for a small water system. G. B. LEONARD *Water Works Eng* 83, 513 (1930) —The water works and the sewerage system in Salisbury, Md., became inadequate in 1924. Investigations of surface and ground supplies showed that shallow wells on low ground produced the only supply worth developing. Six concrete wells have been constructed on the site selected and the Schumaker pond, nearby, can be used as an emergency supply. Aeration, to remove CO_2 , is the only treatment necessary for this water. The water works system is based on an av. daily consumption of 100 gal. per capita. The general layout of the distribution is described. Two pumping operations were required, a low service from the wells to the aerators superimposed over the 600,000 gal. covered concrete storage reservoir, and the other from the reservoir to the mains and to the two 200,000 gal. elevated storage tanks situated in opposite ends of the town. Flue power is used. C. H. BADGER

Utilizing marine gasoline engine as a water works (Oshawa, Canada) auxiliary unit. W. C. SMITH *Water Works Eng* 83, 489-90, 522 (1930) —A detailed description is given of a unique installation used as a standby unit on auxiliary standby plant for a water works system supplying 2 million gals. daily. C. H. BADGER

Building efficiency into an old purification plant. MARTIN E. FLENTJE AND CLARK CRAMER *Water Works Eng* 83, 421-2, 450 (1930) —The reservoir capacity at Lexington, Ky. is 1765 million gals. The water shed area is 12.7 sq. miles. There are 10 wooden tub filters and 2 concrete units with a total rated capacity of 8,325 million gals. daily. Eight addnl. 1 million gals. daily concrete units will be constructed. The reservoir water is fairly constant in chem. compn. and has an av. hardness of 110 p.p.m. The bacteria count is low because of long storage and absence of polluting sources. Microorganisms grow abundantly in the shallow reservoirs in the summer and CuSO_4 treatment is given at least once a week. The sand in the wooden tub filters became badly incrustated because of lack of proper washing facilities. Cleaning with NaOH , Na_2CO_3 , Cl_2 and CuSO_4 solns. was not effective. Since lab. expts. showed that some form of abrasion is necessary for cleaning the sand will be treated with 2% caustic soln. for 24-48 hrs., it should be agitated every 2 hrs. with air piped into the underdrain system. Four samples thus far collected monthly show progressive darkening and lack of luster. The portion sol. in HCl and the loss on ignition have increased. Whether bad incrustation will be prevented remains to be seen. C. H. BADGER

Design and construction of additional filter units. HUGH R. CRAMER AND ARTHUR T. CLARK *Water Works Eng* 83, 450-4 (1930) —The new filter units mentioned (cf. preceding abstr.) are of the general type. Addnl. features are the combining of

the filter effluents, wash water, air wash and rewash through a special fitting. Their construction and operation are described in detail. C. H. BADGER

Building small purification plant at a minimum cost in restricted area. M. P. HATCHER. *Water Works Eng.* 83, 487-8 (1930).—The water supply for Iola, Kan., is obtained from the Neosho River. This river carries much suspended matter and is subject to considerable pollution from a well populated drainage area of 3,000 sq. miles above Iola. About 50 to 75% of the hardness of 100-500 p.p.m. is due to $\text{Ca}(\text{HCO}_3)_2$. Chlorides are usually less than 400 p.p.m. but during low flow periods they run 500-700 p.p.m. because of oil field pollution. The present treatment includes only plain sedimentation and chlorination. In the new treatment plant built on the same site, the raw water is to be pumped to a preliminary mixing chamber where lime will be added. The treated water kept at a high velocity to prevent settling, will then pass to 2 similar reaction tanks for 30 min. agitation with a centrifugal pump and in turn passes to either or both of a 2.5 million gal. and a 3.5 million gal. settling basins for a max. retention period of 3 days. The water is then pumped to the aerator and carbonating chamber. Alum is dry fed to the suction of the auxiliary low service pumps. The reaction proceeds through the carbonating chamber and the alum reaction chamber but the major portion of the floc will be carried to the 3 filter beds, each with a capacity of 2.3 million gal. daily. The filtered water passes in turn through the storage reservoirs situated below the filters, the secondary basins and the pump room. Total storage capacity is 215,000 gal. Cl_2 is applied at the entrance of the last reservoir. C. H. BADGER

Controlling microscopic organisms in public water supplies. FRANK E. HALE. *Water Works Eng.* 83, 353-4, 379-81 (1930).—See C. A. 25, 307. C. H. BADGER

Fluorescence of water in filtered ultra-violet light as an indicator of pollution. H. IVEROVIC. *Z. Hyg. Infektionskrankh.* 112, 54-61 (1931).—The fluorescence of drinking waters was examined in the light from a quartz lamp passing through a black glass filter of NiO , and compared with gelatin standards. Many waters showed a bluish white fluorescence whose intensity was proportional to the consumption of MnO_2 . The org. matter present is considered the cause of the fluorescence. R. B.

Unusual bacterial conditions encountered at the Williamson filtration plant. PAUL C. LAUR. *Proc. 5th Ann. Conference on Water Purification*, W. Va. Univ. Eng. Expt. Sta., *Tech. Bull.* No. 3, 40-8 (1930). *U. S. Pub. Health Eng. Abstracts* 10, W, 115 (Oct. 25, 1930).—Bacterial tests on 5 ten-cc. samples are made daily at the filtration plant. Of 5040 samples of filtered water, 85% gave neg. gas production after 48 hrs. Four types of gas-forming bacteria are described which produce gas but which are not fecal coli. They are thought to originate in the soil or in animal or human feces. Excess chlorination and preammoniation have but little effect on these gas-forming bacteria. C. R. JELLERS

Use of pre-ammoniation at the Springfield, Ill. plant. CHARLES H. SPAULDING. *Water Works Eng.* 83, 514-8 (1930).—The bacterial loads in the raw water have been excessively heavy. Tabulations of lab. expts. using added cultures of *B. coli* with river water show that sterilization with 1 hr. contact was greater when NH_4OH was added just before chlorination but were less satisfactory even with a contact period up to 6 hrs. when lime was also added. After 8 hrs. contact chlorinated samples, with or without lime and NH_4OH , showed practically the same percentage reduction of *B. coli*. Lime by itself is slow in comparison. River water contg. 1.2 p.p.h. phenol is sufficient to cause a chlorophenol taste but NH_4OH added in the portion of $1/2$ lb. NH_4OH to 1 lb. Cl_2 prevented chlorophenol tastes in boiled distd. water contg. 2 to 50 p.p.h. phenol and lime to a pH of 9.6. The packing of the high service pump showed *B. aerogenes* present, estd. at 20 billion per lb. NH_4OH is now added at the rate of 0.7 lb. per million gal. of water at the inlet of the raw water dosing well where the detention is about 30 sec. and Cl_2 is applied at the outlet of the dosing well. The residual Cl_2 as a rule drops by the time the water has passed the filters and about 1 lb. Cl_2 per million gal. of water is sufficient to bring the residual Cl_2 up to 0.2-0.3 p.p.m. Aqua NH_4OH and anhyd. NH_4OH were less expensive to use than NH_4OH salts. C. H. BADGER

Unusual uses of chlorine in a water works system. (Sterilization of new mains, new wells and flooded supplies.) R. W. KERR. *Water Works Eng.* 83, 458-61 (1930).—K. cites the procedures for sterilization by Cl_2 of mains, etc., in New York, Chicago and Indianapolis. The sterilization of newly laid mains in Kansas has not received much attention. The av. time for 9 new supplies to become free from contamination was $6\frac{1}{2}$ months. Gravel used for gravel wells should be thoroughly treated with chlorinated lime water. All turbidity should be pumped from newly constructed wells. Residual Cl_2 was soon shown by the α -tolidine test when chlorinators led several

p p m Cl_2 to flooded supplies in Ottawa, Augusta and Eldorado. The following directions are given for sterilizing small quantities of water: Make a solution of 1 teaspoonful of chlorinated lime in a pint of water, stir 1 table-spoonful of this stock solution into 10 gal. of water and let stand 30 min. With very muddy waters double the dose. To clarify muddy waters add 1 teaspoonful slum to 5 gal. of water. After mixing and settling, the decanted water is sterilized. Boiling the drinking water was found to be impracticable during the Mississippi flood and Cl_2 sterilization was used almost entirely.

C. H. DAPCER

Problems of slow chemical reaction in cold weather. C. T. Hovsen. *Water Works Eng.* 83, 510 (1930).—Lawrence, Kan. obtains its water supply from the Kaw River. The treatment includes plain settling, coagulation, softening, chlorination and filtration. Alk. varies from 125 p p m in summer to 300 p p m in winter and sulfate hardness likewise varies 60 to 125 p p m. The slow reaction of lime in cold water produced the same effect as a reduced retention period in the basins. Settling was not complete even when a dosage of 0.7 grain $\text{Al}_2(\text{SO}_4)_3$ per gal. was applied to the lime basin influent, necessitating a large amt. of wash water for the roughing filters. Attempts during low temp. to lower the alk. to 50 p p m resulted in high carbonates and a heavy ppt. on the final filters, which showed on analysis nearly 60% $\text{Al}(\text{OH})_3$ and the remainder CaCO_3 and MgCO_3 . This ppt. is probably due to the retarded reaction of $\text{Al}_2(\text{SO}_4)_3$ with a highly lime-treated water. Reducing the lime treatment, thereby increasing the alk. to 70-80 p p m, eliminated a large amt. of the ppt. and permitted satisfactory filter operation. In Dec., 1929 bacterial analyses, even with daily increased Cl_2 dosage, began to show poor tests for the coli aerogenes group in the first 1½ miles of mains from the plant, but beyond this distance were neg. This indicated that the reaction of Cl_2 in water at 38-40°F. is slow and that there must be a sufficient contact period before filtration. After unsuccessful results of applying added Cl_2 to the filter influent, 2½ lb. per million gal. of water was then added to the lime basin influent and bacterial analyses showed neg. results. More recent attempts during low temps. to increase the lime dosage to give an alk. of 60-70 p p m caused poor bacterial results immediately. It is evident that increasing the pH value of the treated water decreases the effectiveness of Cl_2 .

C. H. DAPCER

Use of permanganate at the Weirton, W. Va., city water works for the correction of chlorophenol tastes. L. L. LASHLEY. *Proc. 5th Ann. Conference on Water Purification* W. Va. Univ. Eng. Expt. Sta., Tech. Bull. No. 3, 37-6 (1930), U. S. Pub. Health Eng. Abstracts 10, W. 116 (Oct. 25, 1930).— KMnO_4 , 1.45 p p m, applied between the sedimentation basin and filters was entirely successful in removing the phenol tastes and odors and its bactericidal effect was great enough to warrant discontinuance of chlorination.

C. R. FLETCHER

Operating results from well-arranged iron removal plant. C. C. FOURT. *Water Works Eng.* 83, 557-8 (1930), cf. C. A. 24, 6902.

C. H. DAPCER

Arsenic content of the Choussy well water, La Bourboule, and fixation of this arsenic in the organism. R. CHOUSSY, (MULH.) A. COURTONS and CAZALA. *Compt. rend.* 190, 1133-4 (1930).—Arsenic was detected by a modification of the Erlich-Reichardt method, the claim on the HgCl_2 paper being fixed by immersion in a 10% soln. of KI and matched against standard stains prep'd. under the same conditions. A slight excess of KMnO_4 was added to the sample to avoid evolution of H_2S or PH_3 . Arsenic contents of 5.8-6.5 mg. per l. are recorded for the season 1928-29 and analyses of tad. poles kept in the water at the well for 8-10 days showed As contents of 3.25×10^{-4} mg. of As per g. (1.88×10^{-4} mg. per g. in ordinary water). Tadpoles in bottled Choussy water, or in synthetic water of similar compn., contained only 2.73 and 2.325×10^{-4} mg. per g. resp.

R. C. A.

Determination of silicic acid in water. W. SCHERFENS. *Chem. Ztg.* 54, 996-7 (1930).—The colorimetric method depending on the formation of a yellow chromolybdate is described again, preference being given to $\text{K}_2\text{Cr}_2\text{O}_7$ solns. rather than picric acid solns. as standards of color.

T. H. H.

Determination of the "chlorinity" of ocean waters. THOMAS G. THOMAS and RICHARD VAN CLEVE. Intern. Fisheries Commission Univ. Wash. (Seattle), Report No. 3, 5 (1930).—An outline is given of the literature dealing with factors that may produce possible variations between the calcd. and det'd densities. Modifications of Mohr's method, which differ as to type of volumetric app. used, size of sample and concn. of AgNO_3 soln. all give practically the same results. Errors in direct d. detns. were found to result from permitting an elapse of time between collection and analysis of the samples. The detn. of "chlorinity" gives the best method for the calcn. of the

most probably *in situ* at atm. pressure. A table is given for converting "chlorinity" per l to "chlorinity" per kg at 20° C. R. FELLERS

Determining protection of underground pipe through proper survey. P. J. RICHARDS. *Water Works Eng.* 83, 491-2, 525 (1930).—Chem. analyses of the soil and waters or aq. ext. of soil samples taken as required by means of an auger from the surface down to the bottom of the ditch will show the presence and the amt. of all chemicals known to be corrosive, non-corrosive, accelerators or inhibitors of corrosion. The phys. characteristics of the soil affect the rate of percolation of soil waters and will det. the character of the contact between the metal and surrounding materials. Heterogeneity of soils may lead to differences in potential within the soil itself and also between the soil and the buried material. Differences in potential will probably accelerate corrosion proportionate to the current flow. This knowledge will det. whether the pipe will serve without protection or whether an expensive coating is justified. Corrosion surveys have been made in Denver and Pueblo for the use of gas-distributing systems. Strong alkalis are usually destructive. Chlorides and sulfates seem to be the most harmful. Sulfides are potentially dangerous and all soil acids must be considered corrosive. Limestone is not only non-corrosive but seems to be an inhibitor when corrosive chemicals are present. Coatings must be thick enough to protect the metal and wrapped coatings must be used where there may be severe abrasion. It has been estd. that over half of the coatings that have been applied to pipe lines might have been saved if some knowledge of the corrosiveness of the locations had been obtained beforehand and that half of the remainder was either misapplied or underprotected. C. H. B.

The corrosion-checking calcium carbonate-protective layers in water tubes. GERHARD SCHTOKKE. *Z. anorg. Chem.* 44, 40-1 (1931).—The resistance of iron water tubes to rusting is very strongly increased through a CaCO_3 -protective layer. For the formation of the layer, a definite quantity of O must be dissolved in the water. The water must contain a sufficient quantity of dissolved CaCO_3 but must not contain too great an excess of H_2CO_3 . The action of the O may be traced to an electrochem. effect in contrast to the theory of Tillmans, Hirsch and Schilling (*C. A.* 23, 1973). In order to show that the formation of the protective layer is not the result of the adsorption of H_2CO_3 by $\text{Fe}(\text{OH})_2$ but is a cathode reaction, a noble metal was used in place of Fe. Expts. indicate that with suitable arrangements CaCO_3 can be pptd. from a $\text{Ca}(\text{HCO}_3)_2$ soln. on Pt with the aid of O. ALLEN S. SMITH

The continuous control of the salt content of boiler-feed water by means of a visual conductivity measurement. E. ROTHER AND G. JANDER. *Z. anorg. Chem.* 43, 952-4 (1930).—The chloride content of feed water is detd. by leading the water through a specially designed tube contg. Pt electrodes. The detns. are continuously recorded potentiometrically. RUSSELL C. FERN

Problems of a watershed sanitary inspector. S. F. GLASS. *Water Works Eng.* 83, 376 (1930).—Rural sewage-disposal problems and the disposition of dead animals in the Denver, Col., watershed area are briefly discussed. C. H. BADGER

Performance of the Tenafly (N. J.) activated sludge plant. WILHELM RUDOLFS, H. HECKELEKIAN, I. O. LACY AND C. N. HENDERSON. *N. J. Agr. Expt. Sta. Bull.* 502, *Ann. Rept. Dept. Sewage Disposal* 1928-9, 127 (1930).—Chem. and bacteriol. exams. made on the Tenafly activated sludge plant and covering a period of 36 hours' sampling show that a high degree of purification and clarification was obtained. The av. biochem. O demand reduction was 93.5% and the *B. coli* reduction % was 99.99. C. R. FELLERS

Some operating results of the Plainfield (N. J.) disposal plant. WILHELM RUDOLFS AND I. O. LACY. *N. J. Agr. Expt. Sta. Bull.* 502, *Ann. Rept. Dept. Sewage Disposal* 1928-9, 9-12 (1930).—In spite of the staleness of the sewage and the comparatively short detention time, the settling tanks removed 53% of the total suspended solids in addn. to 9-12% suspended solids removed by fine screens. The screened sewage averaged 737 p. p. m. total solids, 256 p. p. m. suspended solids and 190 p. p. m. biochem. O demand. Total reduction in biochem. O demand for the settling tanks and sprinkling filters was 56.2%, $\frac{1}{4}$ of which was taken care of in the sprinkling filters. C. R. FELLERS

Effect of certain chemicals on the vacuum filtration and gravity drying of ripe sludge. WM. H. BAUMGARTNER. *N. J. Agr. Expt. Sta. Bull.* 502, *Ann. Rept. Dept. Sewage Disposal* 1928-9, 17-25 (1930).— FeCl_3 was the most effective treating chemical for ripe sludge. FeCl_3 has a 3-fold effect in sludge drying and filtration: (1) discharge of the dispersed material, coalescence being caused, (2) change of the H ion concn. the isoelec. points where max. pptn. of the colloidal material takes place being thus approached; (3) liberation of CO_2 which makes the material porous and helps

in its flotation. There are peaks in the filtration of FeCl_3 -treated sludge which correspond to definite pH values. These peaks are at a concn of 2 and 7 lb to the cu yd of sludge, resp., FeCl_3 increases the amt of drainage from sludge drying on beds and thus decreases the time for evapn. Treated sludge has a porous nature which allows more surface for evapn. Moisture is removed more rapidly by treating sludge with FeCl_3 . C. R. FELLERS

Certain organic constituents of fresh and ripe sewage sludge. H. HEBKSELEKIAN. N. J. Agr. Expt. Sta., Bull. 502, *Ann. Rept. Dept. Sewage Disposal* 1928-9, 31-6 (1930).—The most abundant constituents of the org. matter of fresh solids are fats and crude proteins. Materials present in smaller percentages are water sol and ale-sol substances, hemicellulose, cellulose and lignin. Fats and cellulose show the greatest decrease in ripe sludge. Resistant substances like waxes and lignin are not decomposed readily and are therefore present in the ripe sludge in larger percentage than in fresh sludge. C. R. FELLERS

Notes on drying fresh sewage solids. ANTHONY J. FISCHER. N. J. Agr. Expt. Sta., Bull. 502, *Ann. Rept. Dept. Sewage Disposal* 1928-9, 29-31 (1930).—Drying of fresh solids in the open is not feasible when treated with lime, alum, or a mixt. of lime and alum on account of poor drainage, objectionable odors, or both. C. R. FELLERS

Sludge drying. Relation between drainage and evaporation. WILHELM RUDOLFS AND I. O. LACY. N. J. Agr. Expt. Sta., Bull. 502, *Ann. Rept. Dept. Sewage Disposal* 1928-9, 26-9 (1930).—Studies dealing with the relation between drainage and evapn. of drying sludge indicate that up to a temp. of 130°F , drainage is a more important factor than evapn. The percentage of total moisture removed, resp., by drainage and by evapn. at various temps. are: 43°F , 91 and 6, 74° , 80 and 14, 100° , 77 and 23, 113° , 60 and 40, and 131° , 47 and 53. The practice of covering sludge beds with greenhouses is beneficial because rain is thus kept off the beds. C. R. FELLERS

Notes on stabilization in a sprinkling filter. WILHELM RUDOLFS. N. J. Agr. Expt. Sta., Bull. 502, *Ann. Rept. Dept. Sewage Disposal* 1928-9, 37.—A summary of chem. results shows the amts. of suspended solids, O consumed, NH_4 , NO_3 , NO_2 , suspended org. N, and sol. org. N present at different depths of the filter beds. In general, the removal of NH_4 nitrogen is well correlated with the reduction on O consumed values. NO_3 and NO_2 values are the only ones which increase with depth. C. R. FELLERS

Determination of pH in the estimation of the productivity of carp ponds. LOUIS ROULE. *Compt. rend. acad. agr. France* 16, 1056-60 (1930).—The pH of the water should be between 6 and 8 and care must be taken to maintain the water within these values. Tests and adjustments of the water should always be made early in the morning and in the shade, as sunlight has a very marked effect on the chem. reactions taking place. J. R. ADAMS

Colorimetric determination of the sulfate ion in water (GUARNIERI) 7. Apparatus for extracting oils from sewage by solvents (U. S. pat. 1,791,398) 27.

BERLINER, J. F. T. The Ammonia-Chlorine Treatment of Water. New York: Natl. Ammonia Co. 22 pp. Reviewed in *Chimie & Industrie* 25, 266 (1931).

PETIT, V. L'eau souterraine. Recherche, captages par sondages. Paris: Ch. Béranger. 130 pp. F. 52. Reviewed in *Chimie & Industrie* 25, 268 (1931).

PRESGOTT, SAMUEL C., and WYNSLOW, C. E. A. Elements of Water Bacteriology, with Special Reference to Sanitary Water Analysis. 5th ed., revised. New York: J. Wiley & Sons Inc. 219 pp. \$2.50.

Vom Wasser, Band IV. Ein Jahrbuch für Wasserchemie und Wasserreinigungstechnik. Edited by FACHGRUPPE FÜR WASSERCHEMIE D. VEREINS DEUTSCHER CHEMIKER. Berlin: Verlag Chemie. 192 pp. M. 17.

Water purification. H. GROECK (to Groeck Wasserveredlung Ges.). Brit. 335,687, Nov. 22, 1928. See Fr. 680,106 (C. A. 24, 3847).

Apparatus for deaerating water. GEORGE H. GIBSON (to Cochran Corp.). U. S. 1,790,911, Feb. 3. Structural features.

Portable water filter. LLOYD E. RABJOHN and RALPH W. HULLENBERG (to Cool Spring Filters, Inc.). U. S. 1,790,947, Feb. 3.

Apparatus for softening water by the lime and soda process. NECKAR WATER-REINIGER MAATSCHAPPIJ. Fr. 693,007, Mar. 29, 1930.

Variable-rate chemical feed pump suitable for water-softening plants. ALBERT B. HODGES (to General Zeolite Co.). U. S. 1,790,708, Feb. 3.

Activated sludge treatment of sewage. JOHN C. DALLAS and MALCOLM WILSON U. S. 1,790,975, Feb. 3 Various details of app are described.

Deodorizing and producing fertilizer from sewage, etc. C. G. WIGLEY and C. PORTS Brit. 335,682, Aug. 12, 1929 Sewage, sludge and the like are treated with a growth of the mycelium of fungi such as *Mucor* or *Rhizopus* developed under forced growth conditions, and under specified conditions of temp., moisture, limited access of air and under subdued light (in a described app.) A material suitable for sale as com. fertilizer may be obtained in 4 days

Apparatus for clarifying waste waters, etc. DEUTSCHE ABWASSER-REINIGUNGSGES. M. B. H., STÄDTEREINIGUNG Fr. 693,040, Mar. 29, 1930.

15—SOILS, FERTILIZERS AND AGRICULTURAL POISONS

J. J. SKINNER AND M. S. ANDERSON

Soils of Connecticut. M. F. MORGAN Conn. Agr. Expt. Sta. Bull. 320, 829-911 (1930) —The ratio of N to org. matter (av. of 100 soils) is 1:20. No serious decline in N content has occurred in the cultivated soils of Conn. during the last 200 yrs. The av. N content of 21 forest soils is 0.205%, which is practically identical with the av. for agricultural soils. The av. percentage of P in the surface 8 in. of Conn. soils is 0.0919, an insufficient amt. for the profitable growing of most farm crops. The availability of the P in many of the soils is very low. K varies from 1.3 to 2.6% on cultivated land with an av. of 1.4%. Nearly all soils respond to potash fertilizers and lime. The av. pH of the soils is 5.3 and the lime requirement per acre approx. 1.3 tons. Numerous greenhouse expts. were conducted to det. the fertilizer requirements of the soils.

C. R. FELLERS

The red soils and Bienhoa of Indochina. M. V. AGAPONOFF *Soil Research* (Supp. to *Proc. Intern. Soc. Soil Sci.*) 2, No. 2, 181-96 (1930), cf. *C. A.* 24, 2822 —The results of chem. and mineralogical studies on samples of soil from Indochina are reported. The soils differ from the basalt from which they are derived in the almost complete removal of Mg, gradual loss of Na, Ca and K, oxidation of FeO to Fe_2O_3 , and content of twice the amt. of Al and Ti as the basalt. The alteration and formation of secondary minerals in the decompos. of basalt are described. The similarities between Bienhoa and laterite are pointed out. The sterility of certain soils could not be accounted for by deficiency of N, P_2O_5 , and K_2O . Soils low in bases in general had poor structure. Data are reported on studies of the absorbing complex of the soils by the method of Gedroiz and on the amt. of SiO_2 and Al_2O_3 sol. in 5% KOH . I. A. D.

Adsorbed bases and unsaturation of Polish sandy soils. W. BUTOWSKI. *Rept. Polish Agr. Expt. Sta.* 2, 195 (1928) —The proportion of adsorbed bases in the soils examd. varied with their mech. compn. and increased with the clay content. The plant covering of forest soils increased the adsorbed bases in the upper horizons. The amt. of adsorbed bases was less in the B-horizon than in the surface but increased at still greater depths approaching the parent rock. Variations in the degree of unsatn. were parallel with those of the base content and were greatest in the loams. The genesis of the soils is discussed. Podsolization is considered to take place in these soils.

B. C. A.

Soils and soil treatments in Quebec. R. R. MCKIBBIN. *Sci. Agr.* 11, 361-8 (1931) —It is the final balance existing in the soil after the addn. of soil treatments that is important, rather than the use of balanced fertilizers. Sweeping conclusions that similar results may be obtained from the use of similar treatments on all light soils or on all heavy soils, are unwarranted. Fundamental soil conditions must be known before effective fertilizer practice can be established. A survey of Quebec's soils is given.

C. R. FELLERS

Soils of Anatolia and eastern Thrace. F. GIESECKE. *Chem. Erde* 4, 551-97 (1930). —A description is given of the various types of soils, with partial mech. and chem. analyses. A soil map of Anatolia is given.

B. C. A.

Nitrate fluctuations in a South Australian soil. J. A. PRESCOTT and G. R. PIPER. *J. Agr. Sci.* 20, 517-31 (1930). —An investigation was conducted to det. the nature and importance of the fluctuations in the nitrate content of the soil during the season under typical conditions of cropping and fallowing. The rate of nitrification is governed primarily by the soil moisture conditions. During the characteristic summer drouth there is little evidence of change in nitrate, the quantities accumulated during the preceding period of activity remaining unchanged. The quantity of nitrate ac-

accumulated during fallowing, as observed over a succession of seasons, appears to approach a const. value of about 20 p. p. m. of the 1st 18 in. of soil, and the greater portion of this remains in the 1st few in. of cultivated soil, sometimes attaining a remarkably high level. It is not washed down until the advent of autumn rains. There is generally no leaching down below the root zone of the cereal crops. At certain periods of the winter the soil moisture may be sufficiently high to result in disappearance of nitrate, when present in quantity, from the surface soil. This depletion is not due to leaching, but possibly to an intake of nitrate by microorganisms. Toward the end of the spring, the rapid nitrate accumulation in the cultivated layer may be due to a possible capillary redistribution caused by rise from the lower layer, or to particularly favorable moisture conditions immediately below the surface of the mulch or in the lower layers of the mulch itself. During the biol. active period of the year the rate of nitrification under fallow conditions is appreciably more rapid than under a crop. Lab. investigations showed that nitrate accumulation was favored by a rather broad range of temp. conditions between 11° and 34°, with no sharp optimum temp. There was, however, a definite optimum moisture content (17%, calcd. on the moist soil), with a rapid fall on the wetter side, leading to nitrate depletion at moisture contents above 19%. Nitrate accumulation occurred with relatively dry soil, although more slowly than under optimum conditions. P. R. DAWSON

The development of the soil profile in North Wales as illustrated by the character of the clay fraction. G. W. ROBINSON. *J. Agr. Sci.* 20, 618-39 (1930).—A study was made of the variations in compn. of the clay fraction in different horizons of some typical North Welsh soil profiles. The changes in the mol. ratio of SiO_2 to sesquioxides ($\text{SiO}_2/(\text{Al}_2\text{O}_3 + \text{Fe}_2\text{O}_3)$) throughout a soil profile afford an indication of the nature of the eluviation which has affected the mineral portion of the soil. The data for the soils studied indicate a general tendency to enrichment of the lower layers in sesquioxides, particularly Fe_2O_3 , at the expense of the surface layers. The SiO_2 -sesquioxide ratio of the clay fraction is an important aid to series definition. The soils of North Wales, particularly in the uplands, have probably been subjected to considerable erosion in the past. The profiles are therefore considered to be more or less truncated. The soil profile appears to be of the brown-earth type, with podsol development in the uplands under heath vegetation. The importance of the parent material is emphasized. North Welsh soils are generally of a loam character. The absence of light sandy soils in situations permitting profile development precludes the occurrence of typical podsoils in the lowlands. The high content of org. matter, resulting from artificial grassland conditions, results in affinities with prairie soils. P. R. DAWSON

The physical and chemical characteristics of certain American peat profiles. IAN V. C. FEUSTEL AND HORACE G. BYERS. *U. S. Dept. Agr. Tech. Bull.* 214, 1-26 (1930).—Based on the data obtained from 38 peat samples, an attempt has been made to distinguish between the various types of peat according to botanical origin of peat material, certain phys. and chem. tests, and response to liming and fertilization. Methods which in many cases are modified standard methods are given for detg. the apparent and true sp. gr., moisture relations, max. shrinkage, reactions, fractional analysis of org. matter, elementary and proximate compn., mineral constituents of the ash, and behavior of certain samples submerged under anaerobic conditions. The max. moisture

grade. O. M. SMITH, Ky Agr Expt Sta., *Research Bull* 308, 447-71 (1930).—The samples were selected from the crops of 1920, 1921 and 1924. Large variations in certain mineral constituents were found in the 2 kinds of tobacco. The extremes for the Burley were 0.70 to 7.0% K, 0.02 to 1.50 Cl and 0.07 to 0.81 sulfate S, while for the dark they were 0.70 to 4.72 K, 0.04 to 2.09 Cl and 0.11 to 0.82 sulfate S. Those samples which had an unusually high chloride or sulfate content did not generally have an abnormal amt of K. In 78% of the comparisons of the Burley tobacco and in 62% of the dark tobacco, more K was found in the good grades than in the common. In 88% of the comparisons of the Burley tobacco and in 48% of the dark tobacco, more Cl was in the good than in the common. In 64% of the comparisons of the Burley tobacco and in 64% of the dark tobacco more sulfate S was found in the good than in the common. The Cl was the same in several comparisons, and the sulfate S in a few. If the 2 Burley crops are combined, the averages of the good grades show 21% more K, 11% more Cl and 1% more sulfate S than the common. A comparison of the Burley and dark crops of tobacco shows that the Burley carries 12% more K than the dark, whereas the latter has 190% more Cl and 11% more sulfate S than the Burley. More K was present in the smoker grades of Burley tobacco than in the filler grades. The K content of the different grades of Burley usually varied in proportion to market quality and price. The same relation existed between the K content, quality and value of the dark tobacco. Org K predominated in the better grades of Burley. Furthermore, the smoker grades had more org K than the fillers. Diseased leaves disturbed the usual percentages of K, Cl and sulfate S present in the leaves. The results indicate that in expts where fertilizers contg K are used on tobacco, particularly of the smoking type, observations should be made of the effects of different amts and combinations of this constituent on the quality as well as the yield of the crop. Bibliography.

C. R. ILLERS.

Impoverishment of forest soils by use of the litter. A. NÉMEC, *Forstarch* 1929, 497-503.—Low proportions of nutrients in many forest soils are attributed to the annual removal of mineral matter in the forest litter. In this respect the available potash is least affected.

B. C. A.

"Single value" soil properties: a study of the significance of certain soil constants. V. On the changes produced in a soil by oven drying. J. R. H. COURTS, *J. Agr. Sci.* 20, 511-5 (1930), cf. C. A. 25, 159.—Results for the loss in wt. of a soil on oven heating can be obtained to a very satisfactory degree of accuracy by an electrically controlled oven. The results obtained by heating soils (brought to equal with atm. at 50% relative humidity) to temps. ranging from 50° to 250° give smooth curves connecting loss in wt. with rise in temp., they lead to the conclusion that there is no sudden alteration in the structure of a soil when it is heated to 100°. The air-dry moisture of a soil, as detd. with sufficient accuracy by the usual methods, is a convenient empirical factor, but not a representation of any fundamental soil property. Expts. show that the losses in wt. on oven drying are greater than are the regains in wt. when the dried samples are brought to equal with an atm. of 50% relative humidity. As the drying temp. is raised, the recoverable loss changes but slightly, while the total loss increases rapidly. Strength is lent the supposition that the recoverable loss is due to removal of free and interstitial water while the unrecoverable loss is due partly to water bound adsorptively or chemically by soil colloids and partly by the destruction of org. colloids. At any given temp. of drying both classes of losses play a role. There is no justification for assuming that there is any sharp line dividing one class of water from the other or that in any process one of the classes of water will be removed while the others are unaffected.

P. R. DAWSON.

The decomposition of urea in soils. T. GIBSON, *J. Agr. Sci.* 20, 549-58 (1930); cf. C. A. 24, 4882.—Investigation of the decompn. of urea in 54 soil samples of extremely varied character showed that the compd. was decompd. readily in all, and very rapidly in most, of the cases. In mountain and heath soils decompn. was generally active, especially in soil tests. Strongly acid peat samples (pH 3.1-3.3) decompd. from 0.44 to 0.86% of their own dry wts. of urea in 21 hrs. at 22-23°. In soln. and soil tests samples from permanent pastures produced a more rapid decompn. than those from cultivated land. In soil tests as high as 1% of the dry wt. of soil was converted to NH_4 in 21 hrs. at 22-23°. Fertile arable soils produced a relatively slow decompn. in soil tests. Forest soils were generally more active than cultivated soils. Strongly alk. soils showed comparatively little activity in soil tests, but greater activity in soln. tests. Although urea is not absorbed by soils, its rapid conversion to NH_4 will generally prevent losses by leaching.

P. R. DAWSON.

Parent material as a factor in soil formation and as a criterion in soil classification.

B. POLYNOV. *Soil Research* (Supp. to *Proc Intern Soc. Soil Sci*) 2, No 2, 165-77 (1930).—The "climatic soil type" is a very broad and general concept. Soils must be further subdivided according to the properties and compn of the rocks from which they are formed, since various rocks differ widely in their influence on soil formation.

I. A. DENISON

A study of the pedogenic processes in an area of lower paleozoic shales. G. H. GETHIN JONES. *J. South Eastern Agr. Coll., Wye, Ann* No 27, 229-47 (1930).—A full account is given of a podsol soil occurring in North Wales. There is a definite acid leaching of the sesquioxides from the eluvial horizon, followed by their subsequent deposition in the illuvial horizon of accumulation immediately below. The nature and extent of some of the pedogenic processes are shown by partial and complete profile chem. analyses of the sedentary soil, down to and including the parent rock. Welsh upland soils are in general very unsatd. with respect to lime because of excessive acid leaching in a humid climate. These soils and parent rock contain a relatively high proportion of MgO , usually about 1%. In the podsolized regions the ratio of exchangeable CaO to org. matter increases with depth, despite the poorer soil dispersion at the lower levels. This shows leaching from the upper layers, without evidence of a CaO horizon of accumulation. A profile study of pH values showed a continuous rise toward neutrality as the relative amt. of unsatd. org. matter decreased with depth. Soil analyses by 48-hr. HCl extn. do not give reliable information as to the abs. chem. compn. of soils differing in their mech. and chem. compn.

K. D. JACOB

The importance of rare elements in the nutrition of plants. O. C. BRYAN. *Proc Assoc. Southern Agr. Workers, 31st Ann. Convention* 1930, 239-42.—A general discussion with particular reference to the use of Cu and Mn salts on the Florida Everglades soils.

K. D. JACOB

Correlations between the specific conductivities of soil extracts, nitric nitrogen and soluble calcium. C. H. WRIGHT. *6th Ann. Bull. Agr. Dept., Nigeria* 1929, 104-9, cf. *C. A.* 25, 761.—Periodical sampling and analysis of cropped soils show that the Ca and nitrate contents of 1:5 soil water exts. are closely correlated with their sp. conductivities. Increased cond. is evidence of biol. activity in the soil.

B. C. A.

Report of the committee on soil reaction measurements. SECOND COMMISSION, INTERNATIONAL SOC. SOIL SCI. *Soil Research* (Supp. to *Proc Intern Soc. Soil Sci*) 2, No 2, 141-52 (1930).—The committee has already published (*Soil Research* 2, 77-139 (1930)) the methods and results of its cooperative investigation on detn. of the pH value of soils. With few exceptions, pH values obtained by 4 independent methods showed satisfactory agreement. Of a total of 23 soils examd., 8 gave erroneous results by the quinhydrone method. It is recommended, therefore, that a preliminary detn. be made to ascertain whether the quinhydrone method is suitable for a given soil. Where the increase in apparent pH value from 10 to 60 sec. is less than 0.2, the method may be considered suitable. The final measurement is made by shaking the soil and water 1 min., adding the quinhydrone shaking again for 10 sec. and measuring the potential after 1 min. In the case of soils considered unsuitable for an exact detn. of pH by this method an approx. detn. can be made on a new sample after 10 sec. In the detn. of pH by the H electrode a sintered glass crucible or funnel is recommended. The colorimetric method did not always agree with electrometric methods, although satisfactory agreement can be obtained with great care.

I. A. DENISON

The relation between pH value and state of saturation of soils. AMAR N. PURI. *Soil Research* (Supp. to *Proc Intern Soc. Soil Sci*) 2, No 2, 181-4 (1930).—The state of satn. of the soil (V) is defined as the ratio of the amt. of absorbed bases in the soil (S) to the amt. of bases the soil is capable of absorbing (T), both values being expressed in equivs. Thus $V = 100S/T$. The value ($T - S$) is detd. by adding 100 cc. 0.2 N $Ba(OH)_2$ to 10 g. soil and 2 g. $CaSO_4$, and titrating after several days against 0.1 N H_2SO_4 with phenolphthalein as indicator. T is detd. by removing all replaceable bases from 10 g. soil with 0.05 N HCl , washing with water and titrating back after adding an excess of $Ba(OH)_2$, as described above. From the values ($T - S$) and T , the S and V values are calcd. A general relation between degree of satn. and pH was noted in the case of 36 soils. The difference in the nature of the colloids present in different soils is shown by fact that the T value per g. of clay varies from 0.252 to 1.516.

I. A. DENISON

Growing onions on the muck soils of New York. J. E. KNOTT. *Cornell Univ. Agr. Expt. Sta., Bull.* 510, 3-34 (1930).—On muck soils that have been under cultivation less than 10 years, increased quantities of N gave a neg. correlation with yield. P had little effect, but a fair response to K was apparent. On new, slightly acid mucks,

1000 lb. of 0-12-18 fertilizer is recommended, while for more acid soils the formula is changed to 3-12-18. C. R. LITTELL

Lysimeter experiments. III. Records for tanks 3 to 12 during the years 1910 to 1924 inclusive. L. C. LAM, J. A. RIZZLER, R. D. WILSON AND J. W. LEECH. Cornell Univ. Agr. Expt. Sta. *Memor.* 134, 5 72(1930) of C. A. 16, 1121. The lysimeters are square concrete tanks with lined shaped bottoms, each containing $1\frac{1}{2}$ bushels of Dunkirk silty clay loam soil. About $\frac{1}{2}$ of the av. 12.5 in. rainfall percolated through the un-cropped tanks, somewhat less than $\frac{1}{2}$ percolated through the cropped soils. About 18% of the water was transpired by the plants on cropped tanks. Applications of lime had no appreciable effect on the proportion of rainfall that percolated through the soil. Liming did not increase the total quantity of N in the drainage water from the tanks in which only non-legumes were grown, nor did it have any effect on the quantity of N in the crops grown on these tanks. In the tanks with legumes, lime produced a decided increase in N both in the drainage water and in the crops. Where legumes were grown in association with non-legumes the quantities of N contained in the latter were nearly as large as when the non-legumes were grown alone. In spite of the space occupied by the legumes and the nutrients they removed from the soil, there was less tendency for the N produced in the crops to decrease from year to year when the rotation contained legumes than when it did not. In the tanks where no legumes were grown, the drainage water alone of the unplanted tanks contained more N than did the drainage water and plants combined from the planted tanks. This is attributed to N-synthesizing organisms which thrive better in the organic matter derived from the plant roots. Tanks in continuous grass had less N than those planted to non-legume crops. The quantity of Ca in the drainage water from the unplanted soil was greater than that in the combined crops and drainage water from the planted tanks. Because of this, an annual conservation of 111 lb. of Ca per acre was effected by cropping the soil instead of leaving it bare. The larger removal of Ca in the drainage water from the unplanted soil than from the planted tanks was due chiefly to the much greater quantity of nitrates leached from the unplanted soil. Liming the soil was not accompanied by an increase in the quantity of Ca in the drainage water or in that of the crops produced. Its loss was decreased by cropping. To replace the yearly loss of Ca from an acre of cropped land requires approx. 250 lb. of CaCO_3 . To recoup the loss from bare soil would require about 900 lb. The quantity of Ca in drainage water was considerably greater where K_2O was applied than where it was not. There was less Mg present than Ca in both the plants and in the drainage. Like Ca its loss was decreased by cropping. Application of lime resulted in a liberation of Mg as indicated by its greater removal in the drainage water. Crops removed more K than did the drainage water. Liming did not increase the K in the drainage water or in the crops. Applications of K_2SO_4 did not result in a larger removal of K in the drainage water or in the crops. The removal of S was much larger by the drainage water than by the crops. Of the S added to the soil in the form of K_2SO_4 , more than $\frac{1}{2}$ was removed in the drainage water. Liming slightly increased the leaching of S in the drainage water. There was never more than a trace of P in any of the tanks. Liming slightly increased the P content of crops from the lined tanks. C. R. LITTELL

Microbiological analysis of soils. T. HAUMGARTEN AND H. BUTENSON. *Landw. Jahrb.* 72, 235 (8(1930)). The authors extend and confirm the results of their previous work (C. A. 24, 6999). By the use of Hesse's medium, cultures taken from soils containing increasing amounts of K fertilizers show increasing development of *Aerobacter chroococcum*, and where heavy dressings of potash have been used there were large numbers of protium. C. OWAN R. GRIFFIN

The value of the simplified Kappen method for determining absorbed bases. D. V. DRUMMOND AND Z. I. STROGANOVA. *Udolnenie i Uroshas (Fertilizers and Crops)* 2, 191-9(1930).—A series of tests was made on various soils comparing the Kappen and Walken Arkman methods for determining the absorbed bases. The methods checked very well. The authors suggest a rapid, approx. method. To 1 part of soil (25, 40 or 50 g.) 2.5 parts of 0.1 N HCl is added, the mixt. is shaken for 1 min., allowed to stand 5 min., filtered and an aliquot taken and titrated with methyl orange as indicator. Tests with 16 different soils showed that the method checked very well with those of K. and of B-A. Just as in the other methods the new method works best on soils free from lime and Ca phosphate and with a low content of absorbed bases. J. S. J.

The antimony electrode. M. S. DU TOIT. *South African J. Sci.* 27, 227-35 (1930).—The Sb and H₂ electrodes were compared in a large no. of soils and found to agree closely. In the potentiometric titration of Fe and Al salts the Sb electrode proved highly satisfactory. Its temp. coeff. was found to be 0.00113 v. per degree in

soils of varying pH values. I rankle and Wilman's earlier formula was confirmed. The use of the H_2 electrode was advocated for soil surveys. C. E. JEFFREYS.

An approximate method of mechanical analysis of soils for field purposes. P. KAMPMAN. *Soil Research Group to Proc Intern Soc Soil Sci* 2, No 2, 152-64 (1930). — A rapid method is described. The soil is dispersed by means of a mortar and rubber pestle after treatment with H_2O_2 . In the case of certain soils the sample is pretreated with 0.2 N HCl or 6% H_2O_2 . After a suitable period of settling the clay fraction is estimated by drying its supernatant, and the sand fraction by decantation and measurement of its volume. I. A. DENISON.

Neubauer analyses and field experiments. OBERDORFFER. *Superphosphat* 6, 83-6 (1930). — In 28 field experiments on deep loam soils the crop responses to potash and phosphate fertilizers generally were in good agreement with the potash and phosphate deficiencies in the soils as indicated by Neubauer tests. K. D. JACOB.

The methods of determining the phosphorus requirement of the soil. A. L. MASLOVA AND O. M. DOPRYVOROSKAYA. *Udobreniia Urozhai (Fertilizers and Crops)* 2, 276-81 (1930). — Soils from 27 field experiments were tested out in the lab. by the method of titration curves and soly of P_2O_5 . Into 100-cc flasks 2, 4, 6, 8 and 12 cc of 0.1 N HCl were introduced, 10 g of soil was added and the whole made up to vol. After 24 hrs the filtrates were analyzed for pH and P_2O_5 . It was found that for 2 different soils the soly of P_2O_5 with HCl might be the same, but the pH at which the same quantity of P_2O_5 is dissolved differs. Thus in chernozem soils the pH value varies from 3.3 to 4.45, whereas in the dark gray loams it is 2.57. A comparison of the crop yields from the experimental plots with the titration curves shows a certain correlation between the pH and response to P fertilization. J. S. JOFFE.

Testing soil for available phosphoric acid by the Winogradsky method. E. M. NURBAIS. *Fertilizer Green Book* 12, No 2, 39-41 (1931). — Directions are given for carrying out the *Azotobacter* test for phosphate deficiencies in soils. K. D. JACOB.

Phosphoric acid and plant growth. M. P. MORTON. *Calif Dept Agr, Monthly Bull* 19, 766-8 (1930). — A general discussion. Most Calif soils are deficient in P .

Phosphorus-fixing compound in the soil. A. H. MEYER. *Science* 71, 461 (1930). — An Fe compound existing as concretions in southern soils and responsible for the fixation of P has been discovered. The P is present as a basic ferrous phosphate of very low soly. B. C. A.

Influence of phosphoric acid on the cropping power of seed potatoes. DENSON. *Superphosphat* 6, 120-4 (1930). — Phosphate deficiency in soil not only reduces the crop yield of potatoes but lowers the seed value of the tubers, which in the subsequent season produce smaller crops of lower starch content. Moreover, the smaller crops in the 2nd season are not improved by phosphate fertilizing. B. C. A.

Critical consideration of fertilizer practice and crop yields in German agriculture. ALMEYER. *Superphosphat* 6, 115-20 (1930). B. C. A.

Chemistry and the fertilization of Italian soil. N. PARRAVANO. *Industria chimica* 5, 1333-46 (1930). — A history of the development of the fertilizer industry in Italy.

Factors in the production of synthetic and natural nitrates. I. E. VANDERKINDEN. *Fertilizer Green Book* 12, No 2, 15-6 (1931). K. D. JACOB.

Developments in the production and use of concentrated fertilizers. C. H. KUNSMAN. *Fertilizer Green Book* 12, No 2, 20-2 (1931). K. D. JACOB.

Modern methods in fertilizer manufacture. B. LESLIE. *PMSLIE Sci Agr* 11, 205 (1931). — A general discussion. C. R. FELLERS.

New fertilizer mixer calculator. E. L. RAYNAULD. *Sci Agr* 11, 159-61 (1930). — A device is described. It is simple, inexpensive and fairly accurate. C. R. FELLERS.

Development of knowledge of rational application of commercial fertilizers and new knowledge in this field. O. FYGIEL. *Kunststnger u. Leim* 27, 331-6 (1930). — An account of the chem. methods and vegetation tests employed for estimating the fertilizer needs of the soil. C. J. SCHOLLENBERGER.

Possibilities of sulfur as a soil amendment. G. S. FRANKS. *Tex Agr Expt Sta., Bull* 414, 3-56 (1930). — Field and pot experiments were conducted over a 4 yr period. Chem. analyses show that alfalfa, cabbage, cotton, onions and turnips take up much larger quantities of S than corn, rice, oats and wheat. Some Texas soils are low in S . S is brought down by rain and also is supplied by irrigation water and in most corn fertilizers. The amt. brought down by rain in Texas averages 4 to 12 lb a year on each acre, varying with different sections. Pot experiments show that S alone gave very poor results, but when it was used to supplement a complete fertilizer in pots watered with distilled water which

contained no S, it increased the yield of crops (in some cases). Adds of S did not increase the amts. of N or K taken up by crops in pot expts., although they increased the S taken up and slightly increased the P. There was a tendency for the S removed by crops to increase as the S content of the soil increased. Oxidation of S had practically no effect upon the active P or active K in the soils tested, but increased the permeability of some of the soils to water. S is not recommended as a fertilizer on soils in Texas, since a sufficient amt. of S is present in the soils, or is supplied by rain or irrigation water or by com. fertilizers. S or CaSO_4 may be recommended in special cases on soils which run together under irrigation, or which contain black alkali. It is possible that the use of coned com. fertilizers contg. little or no S may cause a deficiency of S in soils in some sections of the country, especially for crops which require comparatively large amts. of S, such as alfalfa, cotton, cabbage and onions. C. R. FELLERS

The effect of the admixture of magnesium with lime applications. O. K. ZIKHMAN, KEDROV. *Udobrenie i Urekaia (Fertilizers and Crops)* 2, 186-9 (1930).— CaO , CaCO_3 (C. P.) and two kinds of limestone, one contg. 47% CaO and 23.3% MgO , the other 44.7% CaO and 25.2% MgO , were compared on pot soil. The results showed that the Mg limestone was even slightly more effective than the pure CaCO_3 . Another series of expts. was conducted with pure CaO , MgO , CaCO_3 and MgCO_3 alone or in combinations. Again the MgCO_3 showed no injurious effects, and in combination with CaCO_3 it stimulated the effectiveness of the latter. On some field plots the amt. of lime applied was almost equal to the hydrolytic acidity, and in every case the Mg-contg. limestones were very effective. It is concluded that limestones contg. CaO and MgO in the ratio of 2:1 or even 1:1 are not injurious, on the contrary it might be more beneficial than pure CaCO_3 . J. S. JORJE

The repressive effect of lime and magnesia upon soil and subsoil potash. W. H. MACINTYRE, W. M. SHAW AND J. B. YOUNG. *J. Agr. Sci.* 20, 499-510 (1930).—Data are presented from 5 lysimeter studies conducted over a 15-yr. period at the Univ. of Tenn. Agr. Expt. Sta. A 12-yr. expt. shows that economic adds of CaO , MgO , limestone and dolomite, incorporated throughout the soil, depressed the soly. of native supplies of K. One-ton and 32-ton adds. of CaO and MgO produced the same repressive effect, even when supplemented by excessive quantities (added at once) of sulfates of Ca and Mg from FeSO_4 , and progressive increments of the same salts from pyrite and S. A 15-yr. study with excessive quantities of 7 forms of Ca and Mg showed a decided decrease in the soly. of K in the surface soil and a marked decrease in the subsoil as a result of the influx of bicarbonate-impregnated percolates. A 4-yr. study with "light" and "heavy" forms of MgO and MgCO_3 and cryst. $\text{MgCO}_3 \cdot 3\text{H}_2\text{O}$ showed the same consistent depression in K soly. Surface zone incorporations of Ca(OH)_2 limestone and dolomite produced neutral-salt impregnated percolates that gave some indication of K liberation in an underlying unlimed zone of surface soil. Sub-surface zone incorporations produced the same repressive effect as the full depth incorporations. When K adds. were made by means of red clover hay along with Ca(OH)_2 limestone and dolomite, a decided decrease in the outgo of the K resulted. The results demonstrated that the liming of rock-derived soils under humid conditions will depress the hydrolytic disintegration of both the original K complex and that formed by fixation of added sol. K salts. It is further shown that the protective, or buffering, effect becomes more pronounced with increase in concn. of the bicarbonates of Ca and Mg. P. R. DAWSON

Plant-food value of mixed and fertilizing materials. BRUNO WÄSSER. *Metall.-Börsen* 20, 2021-2, 2077-8 (1930).—A table shows ratios of N, P_2O_5 and K_2O contents of a number of common plants grown in Germany. The mean ratio for some 42 plants is $\text{N}:\text{P}_2\text{O}_5:\text{K}_2\text{O} = 1.047:1.42$. A similar table shows the same constituents in natural manures. The mean ratio here is $1.046:1.52$. By combining these facts with the av. amts. of N, P_2O_5 and K_2O removed by various crops from the soil, a fair idea can be gained in regard to fertilizer requirements in Germany. The ratio of fertilizers actually used in Germany, in the case of 12 of the most important crops is $\text{N}:\text{P}_2\text{O}_5:\text{K}_2\text{O}$ lime = $1.049:1.256:0.66$. Approx. the same ratio is found in artificial mixed fertilizers, such as KNO_3 , $\text{N}:\text{K}_2\text{O} = 1.156$, Nitrophoska, $\text{N}:\text{P}_2\text{O}_5:\text{K}_2\text{O} = 1.0825:1.413$; Am-Snper, $\text{N}:\text{P}_2\text{O}_5 = 1:1.55$, Am-Snp-Ka, $\text{N}:\text{P}_2\text{O}_5:\text{K}_2\text{O} = 1.135:1.53$. A table shows the ratio of fertilizers applied to soil in some other European countries, also in Egypt and Japan. The av. ratio in these countries is $\text{N}:\text{P}_2\text{O}_5:\text{K}_2\text{O} = 1.2:1.1:0.51$. For the U. S., estimates are given for 3 ratio-types, $\text{N}:\text{P}_2\text{O}_5:\text{K}_2\text{O} = 1.333:1; 1.5:1; 1.3:1$. S. L. MADORSKY

Recent experiments on the preparation of organic manure. GILBERT J. FOWLER. *Agr. J. India* 25, 383-85 (1930).—Fermentation of org. matter in the prepa. of artificial

manure is facilitated by chopping the material into small pieces. The duration of fermentation is greatly diminished by the preliminary building up of a vigorously fermenting mass of material to serve as "activator," with which smaller quantities of less completely fermented materials are systematically mixed at suitable intervals. Insect larvae are practically eliminated by completing the fermentation anaerobically in pits. The necessary fermentation organisms may be obtained from either cow dung, cattle urine or night soil, the latter always producing the most active fermentation fungi. Lungs play an important part in the fermentation process, and materials containing a high percentage of lignin cannot be rapidly fermented. There is some evidence that N fixation occurs during the process of fermentation. Data on the prepn of artificial manure from prickly pear, leaves, town refuse, mahua flowers, banana waste, leguminous plants and shrubs, sunn hemp (*Crotalaria juncea*), weeds and green pea stalks are tabulated.

K. D. JACOBS

Fermentation of compost and liquid manure. FRANZ SCHACHT. *Kunstdünger u. Leim* 27, 30*-8(1930).—Excessive fermentation induced by pumping liquid manure over solid stall manure can add nothing to the fertilizer value of the compost, the only advantage is in reduction of vol. by evapn., but this is accompanied by loss of N by volatilization as $(\text{NH}_4)_2\text{CO}_3$. Addn. of 18% superphosphate at the rate of 8-12 kg./cu m. compost prevents loss by fixing NH_3 as nonvolatile phosphate. C. J. S.

Factors that influence nitrogen fixation in soils. THOMAS F. MANUS. *Del. Agr. Expt. Sta. Bull.* 167, *Ann. Rept. Director* 44-52(1930).—Twenty years' expts. (with a rotation of wheat, mixed hay, corn and soy beans) are reported on N fixation in limed and unlimed plots receiving various fertilizer treatments. Conclusions: The use of lime, unless accompanied by K, resulted in less N gain in 20 years than shown by the unlimed plots, and lime was not a marked factor in N fixation. N losses from percolation in Del. probably reach 20% of that removed in the crops. K and P in combination were the most active factors in stimulating N fixation. K was the most active single element and the only mineral element that the crops removed in excess of that applied. C. R. FELLERS.

The absorption of ammonium and nitrate nitrogen by various plants at different stages of growth. JAMES A. NAPTEL. *J. Am. Soc. Agron.* 23, 142-58(1931).—Ammonium and nitrate N are absorbed by seedlings in appreciable amts. throughout all stages. $(\text{NH}_4)_2\text{SO}_4$ and $\text{Ca}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O}$ were used as sources of the N and similar results were obtained with all 3 plants (cotton, wheat and corn). The ammonium N is absorbed more rapidly in the early stages, after which the reverse is true. The best growth and earliest fruiting are obtained when there are both forms of N present. The easily sol. N of the seed and sap at the time of sprouting is present practically entirely in the ammonium form, and the N in this form is utilized much more rapidly than the nitrate N. The absorption of ammonium N increased as the acidity of the culture soil decreased, while the absorption of the nitrate N was very slightly affected. The highest total N absorption usually occurred at pH 6.0 and when both forms of N were present. A theoretical discussion of the observed phenomenon is included. J. R. ADAMS.

A comparison of various forms of nitrogen fertilizers. E. I. RAYNER. *Udobrenia i Urazhas (Fertilizers and Crops)* 2, 291-300(1930), cf. *C. A.* 24, 1923.—The fertilizer values of NaNO_3 , $(\text{NH}_4)_2\text{SO}_4$, $\text{CO}(\text{NH}_2)_2$, CaCN_2 and NH_4HCO_3 were tested in pot expts. on podsolized sandy loam, loam, degraded chernozem, deep chernozem, dark chestnut brown, Turkestan loess and red soils (Roterde) with oats as the exptl. plant. None of the forms of N proved superior under all conditions. In the unsatd. soils cyanamide was superior to any form of N. On heavy loams $(\text{NH}_4)_2\text{SO}_4$ showed slightly higher effects. It is of interest to note that wherever $(\text{NH}_4)_2\text{SO}_4$ was not efficient the cyanamide was. It might be correlated with the different effects of these 2 forms of N on the reaction of the soil as shown in the tables presented. On some of the podsolized soils the tests were repeated, but to one series lime was added. The $(\text{NH}_4)_2\text{SO}_4$ was as good as the other forms of N on the limed soils. Cyanamide was superior to the others, even with lime, provided acid phosphate was used as the source of P. With pptd. phosphate the cyanamide was inferior. On freshly limed sandy loams (podsoils) cyanamide was inferior, but on podsolized loams it was very efficient. Tests were also conducted with various moisture contents, but the results were of no significance. J. S. JOYNS.

Effect of fertility on the carbohydrate-nitrogen relation in the soy bean. F. A. WELTON and V. H. MORRIS. *Plant Physiology* 5, 607-12(1930).—Soy beans grown in sand (3 parts of sand to 1 of Wooster silt loam) contained more dry matter and more total carbohydrates than did those grown in either soil (Wooster silt loam) or manure (3 parts of manure to 1 of Wooster silt loam). In general, the increase was due chiefly

to easily hydrolyzable carbohydrates, cellulose and lignin. The increase in carbohydrates in the plants grown in the sand was not accompanied by a simultaneous decrease in N as has been found with certain non legumes (C A 14, 1725-6). The high N content of these plants, however, was associated with the development of relatively large nos. of nodules on the roots of the plants. The stems of the plants grown in the sand were comparatively tough and rigid and not inclined to lodge. WALTER THOMAS

"Kalkammonsalpeter." O. NOLTE AND H. MCENZFAC. *Mitt. deut. Landw. Ges.* 45, 37(1930).—Comparison is made of the efficiency of "Kalkammonsalpeter," $(\text{NH}_4)_2\text{SO}_4$ and NaNO_3 in field trials. In dry seasons NaNO_3 was particularly effective, but in wet summers, when leaching losses were considerable, $(\text{NH}_4)_2\text{SO}_4$ proved the more profitable. "Kalkammonsalpeter" is preferable to either of the above for cereals and root crops. B. C. A.

The production of superphosphate from Khibinsk apatite. S. I. VOLKOVICH, L. I. BERLIN AND L. B. GRINSHPAN. *Udobrenie i Urozhai (Fertilizers and Crops)* 2, 300-12 (1930), cf. C A 25, 763.—A rich apatite obtained from nepheline apatite rock by fine grinding and sieving or flotation is used on a semifactory scale for the production of acid phosphate. The quantity of H_2SO_4 to be used was based on the following equation: $2\text{Ca}_5(\text{PO}_4)_3 + 7\text{H}_2\text{SO}_4 + 17\text{H}_2\text{O} = 3\text{CaH}_2(\text{PO}_4)_2 + 11\text{H}_2\text{O} + 7\text{CaSO}_4 + 2\text{H}_2\text{O} + 2\text{H}_2\text{O}$. It was found that by using 83% of the rock milled to less than 0.1 mm a 40% P_2O_5 raw product was converted to acid phosphate of a good phys. condition with 18% sol. P_2O_5 . With 83% milled to less than 0.1 mm the raw product contained 38.9% P_2O_5 and the final acid phosphate 10.2% P_2O_5 . When the quantity of particles less than 0.1 mm fell below 23-34% P_2O_5 the phys. condition of the final product was poor. It was found possible to mix the high grade nepheline apatite with the Vyatka raw phosphate low in P_2O_5 (24.5% P_2O_5), half and half and still obtain a good product. J. S. JORRE

Fertilization trials with potassium ammonium superphosphate on root crop. DENSCHE. *Superphosphat* 5, 53(1929).—The efficiency of the N of "Kali-ammon-superphosphat" is similar to that of the simpler nitrogenous materials, and the fertilizer proved suitable for use on acid soils and for acid sensitive plants. B. C. A.

Changes in soil reaction produced by ammonium, potash and potash ammonium superphosphate. GERLACH. *Superphosphat* 5, 282(1929), cf. C A 25, 164.—The increasing acidity of many German soils is more definitely attributable to the normal processes of cropping and leaching than to the use of physiologically acid fertilizers. Systematic liming or marling is preferable to the withholding of valuable, if slightly acid, fertilizers. B. C. A.

Composition of ammoniated superphosphate. K. D. JACON. *Phosphorus Digest*, pp. 5-6 (Nov., 1930), *Fertilizer Green Book* 12, No. 2, 10-1(1931).—The citrate-insol. P_2O_5 formed when superphosphate is treated with relatively large quantities of NH_3 is present principally as $\text{Ca}_3(\text{PO}_4)_2$. The phys. and chem. properties of $\text{Ca}_3(\text{PO}_4)_2$ are quite different from those of the P compds. present in phosphate rock, which is composed principally of Ca fluophosphate having essentially the same constitution as fluorapatite, $3\text{Ca}_5(\text{PO}_4)_3\text{CaF}_2$. Recent expts. indicate that $\text{Ca}_3(\text{PO}_4)_2$ has approx. 75 to 80% of the fertilizer value of mono- and di. Ca phosphates during the first growing season, while phosphate rock usually has a much lower value. K. D. JACON

The use of superphosphate on acid soils. H. KAPPEN. *Superphosphat* 6, 66-8(1930); cf. C A 24, 4576.—The long-continued use of superphosphate on heavy clay and light sandy acid soils does not have a significant effect upon the pH values of such soils. When used continuously on limed soils superphosphate tends to further increase the pH values of the soils. Basic slag and Rhenania phosphate have approx. the same value in reducing soil acidity, and it seems that with the quantities ordinarily used in actual practice neither of these materials reduces soil acidity to a significant extent. K. D. JACON

Soil fertilization for sugar beets. JAMES TYSON AND M. M. MCCOOL. Mich. Agr. Expt. Sta., *Special Bull.* 205, 3-31(1930).—Field expts. conducted on 3 of the most important sugar-beet soils in Mich. show the best fertilizer ratios to be 1-4-1 or 1-4-2. From 400 to 600 lb. of 4-16-4 or 4-16-8 mixts. gave most profitable returns. NaNO_3 when applied all at one time to the prep. soil before sowing the seed was fully as effective as when the same amt. was applied in instalments. C. R. FELLERS

Fertilization and crop quality in root crops. KLEBERGER. *Superphosphat* 5, 74-6 (1929).—The effect of unbalanced fertilization on the quality of sugar beet and potatoes is examined. In this respect nitrogenous fertilizers are of first importance. The use of phosphate fertilizers in amts. based on Neubauer trials may be unsatisfactory, since this method of examn. allows of no consideration of quality. There is no foundation for the opinion that phosphate fertilizers are of minor importance for potatoes. B. C. A.

Action of ammonium sulfate and of sodium nitrate on the yield and starch content of potatoes. O. ENGELS *Fortschritte Landw.* 5, 97(1929);— $(\text{NH}_4)_2\text{SO}_4$ produced higher crop increases and starch content of potatoes than did NaNO_3 . With each fertilizer there was an economic return. B. C. A.

Sources of nitrogen for potato fertilizers in Aroostook County. B. L. BROWN, P. V. OWEN AND E. R. TOFFY *Maine Agr. Expt. Sta. Bull.* 354, 39 pp (1930).—During 16 years' exptl. work $(\text{NH}_4)_2\text{SO}_4$ produced a higher av. yield than NaNO_3 , but comparable results depend primarily upon seasonal conditions. The most important effects of seasonal variations are (1) leaching of NaNO_3 during periods of heavy rainfall and (2) unavailability of $(\text{NH}_4)_2\text{SO}_4$ during periods of drought, especially on very acid soils. Although there is very little difference in the yields of potatoes following the use of inorg. sources of N as compared with inorg. N supplemented with org. N, it is not recommended that the practice of adding some org. N be discontinued, though the amt. may be reduced. From 5–6% of NH_3 (4–5% N) in the potato fertilizer mixt. gave the best results on Caribou loam soil type. N in potato fertilizers is an absolute necessity for high yields. Leunassalpetar, NH_4NO_3 , NH_4Cl , $(\text{NH}_4)_3\text{PO}_4$, $\text{Ca}(\text{NO}_3)_2$ and especially urea compared favorably with $(\text{NH}_4)_2\text{SO}_4$ and NaNO_3 as single sources of N. Good results were also obtained with CaCN_2 , when the amt. did not exceed 60 lb. per ton of mixed fertilizer. The degree of acidity and alkali of the soil and its ability to hold moisture are important factors in detg. the effectiveness of various N materials. C. R. FELLERS.

Effect of nitrogenous fertilizers on pastures. O. NOLTE, H. MCNEAG AND H. KOCI *Mitt. deut. Landw. Ges.* 44, 385(1929).—Results of meat and milk trials on fertilized pastures are recorded. B. C. A.

Effect of various fertilizer treatments on the yield and chemical nature of permanent pasture sod in the Piedmont section. J. P. LAMASTER *Proc. Assoc. Southern Agr. Workers, 31st Ann. Convention*, 148–58(1930).—Application of either a complete fertilizer or a N fertilizer alone greatly stimulated pasture growth for a period of 2 to 3 months following the application. A second and third application of a N fertilizer during the growing season greatly stimulated growth for a few weeks, the effect rapidly diminishing as the season advanced. Potash and phosphate fertilizers alone had very little effect in promoting growth of pasture, and little difference in the rate of growth was noted on the limed and unlimed plots. When superphosphate and lime were applied at the same time, reduced yields of green material and total P_2O_5 were obtained. On the basis of the whole season's growth, no significant differences in the percentages of N in the dry matter resulted from the various fertilizer treatments. The Ca content of the grass on all plots was very similar and followed the same trend through the season on all plots, the percentages being highest during the early season and lowest during the latter part of the season. The P content of the grass on all plots was also very similar. The percentages of P were high during the early growth period, low during midsummer and high again during the latter part of the growing period. E. D. J.

Fertilizer work with cabbage. L. M. WAIN *Miss. Agr. Expt. Sta. Circ.* 91, 2–4 (1930).—Two yrs' field expts. showed N was the only element which when added alone gave an increase in yield. The largest increase and net profit were obtained by the use of 532 lb. of NaNO_3 , 1250 lb. of superphosphate and 100 lb. of KCl per acre. C. R. FELLERS.

Effects of certain soil conditions on the yield and quality of Burley tobacco. C. A. MOORE *Tenn. Agr. Expt. Sta. Circ.* 33, 4 pp (1930).—Tenn. tobacco fertilizers show an av. content of 2.5% N, 9% P_2O_5 and 4% K_2O . Applications of from 500 to 1000 lb. per acre were more profitable and gave higher yields than light applications of from 100 to 450 lb. per acre. The use of manure also gave increased yields and profits. C. R. FELLERS.

Long-time fertilizer experiments (with peaches) in northeast Georgia. H. M. MCKAY *Proc. Assoc. Southern Agr. Workers, 31st Ann. Convention*, 318–25(1930).—Over a period of 10 years N-K fertilizers gave the greatest total wt. of fruit and the greatest no. of fruits per tree. Applications of P fertilizers alone gave poorer results than those obtained on the check plots. K fertilizers increased the yield of fruit, but to a less extent than N fertilizers. For the period of the expt. the increase in the circumference of the trees was 33 to 50% greater when N was used in the fertilizer, and the terminal growth was also greater. P hastened the maturity of the fruit while N retarded maturity to a marked extent. K. D. JACOB.

Cotton fertilizer experiments, 1930. Sources of nitrogen, supplements and time and method of application. G. A. HALE *Ga. Agr. Expt. Sta. Circ.* 91, 4 pp (1930).—Field expts. showed that the P in superphosphate or treble superphosphate (43%

P_2O_5) was somewhat more efficient than that in $(NH_4)H_2PO_4$ or $NH_4H_2PO_4$. A small amt. of cottonseed meal used with $NaNO_3$ caused the production of more cotton than the $NaNO_3$ alone. The ordinary complete fertilizer contg. N, P and K were improved by the addn. of small amts. of Mg and Ca. C. R. FILLERS

Chemistry and plant protection. A. CHWALA. *Chem. Ztg.* 32, No. 6, 45-8 (1930). C. reviews the chem. insecticide and fungicide materials, their classification, application and properties. He criticizes the crude methods used in chem. labs. for detg. the toxic effectiveness of these materials and points out the important influence of phys. factors, such as soly., dispersum, size of particles, surface tension, resistance to weather, wetability, etc. on the toxicity. S. L. MACOSKEY

Ability of dry fungicides to adhere to seeds. A. SHUKIN. *Udarnik's Uralskii (Fertilizers and crops)* 2, 26, 12 (1930). Into glass stoppered jars (100 cc. capacity) 20 g. of grain is introduced and placed in cups of a revolving drum. The fungicide is added and the drum kept at a definite speed for different time periods. The seeds are taken out, passed over a 10-mesh sieve (10 open sq. per square cm.) and washed, and the amount of dusting material absorbed is detd. Wheat, oats and barley were used. Ca arsenate, Paris green, $CuCO_3$ and Cr_2O_3 dusts were washed with a weak soln. of HCl. It was found that $CuCO_3$ and anhyd. Cr_2O_3 adhere best to wheat and Ca arsenate to oats. The longer the seeds are rinsed with the dusts the more is absorbed. J. S. JORR

Review of research on the control of wireworms. C. A. THOMAS. Penn. Agr. Expt. Sta., Tech. Bull. 259, 232 (1930). One of the most effective baits consists of 6 lb. of $Ca(CN)_2$ per 1000 ft. of crop row. The rows should previously be liberally baited with wheat, oats or corn so as to attract the wireworms. The use of CS_2 emulsions on limited areas as in greenhouses or gardens is also fairly effective. There are 12 pp. of bibliography, many of the references being briefly reviewed in the text. C. R. FILLERS

Possibility of a new insecticide for use on citrus. L. B. RILEY AND G. A. HERRBURN. *Farming in S. Africa* 5, 20, 241 (1930), cf. C. 4, 24, 3591— $NaSiF_6$ has for the Natal fruit fly a killing power 10 times that of Pb arsenate. Preliminary expts. with strengths needed to kill the fly caused no damage to citrus foliage, but it is known that considerably higher concns. are capable of severe foliage burning, and the limit of concn. for safe spraying has not yet been detd. K. D. JACOB

The influence of antiseptics on the supply of nutrients in the soil. G. A. CHIGAREV AND N. B. MYAKIN. *Udarnik's Uralskii (Fertilizers and crops)* 2, 312-7 (1930).—By using CS_2 or polychlorides (a mixt. of chlorinated hydrocarbons of the benzene series, b. 130-200°), consisting primarily of the isomers of dichlorobenzene with some mono- and polychloro derivs.) as soil disinfectants it was found that the nitrates are inhibited to a great extent by the CS_2 , but are stimulated by the polychlorides. NH_4 is stimulated by both disinfectants. Very little effect was noted on the P regime in the soil. J. S. JORR

Mango hoppers and mildew and their control. P. V. WAGLE. *Panna Agr. Coll. Mag.* 21, 170-6, (1929).—Mildew on mangoes is effectively controlled by dusting with S alone or with a mixt. of 6 parts of S and 1 part of $Ca(CN)_2$. S checks the activities of the mango hopper, but better control is obtained by dusting with the S-sewage mixt. K. D. JACOB

The economics of pyrethrum. JOHN GLASSFORD. *J. Econ. Entomol.* 23, 574-7 (1930).—Japan now produces 9% of the world's pyrethrum flowers. Imports into the U. S. increased from 3 million lb. in 1923 to 9 million lb. in 1929. Average annual invoice values decreased from 47c per lb. in 1923 to 15c per lb. in 1929. A pyrethrum spray contg. 0.04% of pyrethrins oleoresin and activated with soap costs 2.2c per gallon compared with 0.3c for a 1% lubricating oil spray, 1.2c for a Pb arsenate spray, 1.75c for a lime-sulfur spray and 1.84c for a nicotine spray contg. 0.04% alkaloid. C. H. RICHARDSON

Experiments with insecticides against cattle grubs (*Hypoderma* spp.). F. C. BISHOPPE, E. W. LAKE, R. W. WELLS AND H. S. PETERS. *J. Econ. Entomol.* 23, 832-63 (1930).—Single and repeated applications of various insecticides were made against cattle grubs in the backs of cattle. Special attention was given to insecticides in dust form. Ground derris root and dust carriers contg. derris ext. gave excellent results under varying conditions. Tobacco powder and dust contg. free nicotine and nicotine sulfate also gave a high degree of control. C. H. RICHARDSON

A comparison of the toxicities of *p*-dichlorobenzene and naphthalene to the confused flour beetle (*Tribolium confusum* Dur.) (Coleoptera). RUSSELL S. LEHMAN. *J. Econ. Entomol.* 23, 935-40 (1930).—Adults of *T. confusum* were exposed to a series

of comens of *p*-dichlorobenzene and naphthalene at 30° and 60-65% relative humidity. Air said with *p*-dichlorobenzene is strongly anesthetic to this insect. Based upon the time required to kill 50% of the insects, naphthalene was 10-20 times more toxic than *p*-dichlorobenzene at the same comen. C H RICHARDSON

Arsenical sometimes injures peach trees. W C DUTTON Mich Agr Expt Sta. Quart Bull 13, 55-6 (1930)—1 lb arsenate used alone or in combination with following materials in the form of a spray or dust during the growing season may cause severe injury to leaves, wood and bark, often resulting in defoliation. Light applications and dil comens of sprays and dusts, as well as mixing of the 1 lb arsenate with lime, are recommended as precautionary measures. C R FELLERS

A new fungicide for the control of both peach and apple diseases. L A NIVEN Free Assoc Southern Agr Workers, 31st Ann Convention, 308-10 (1930)—With the exception of little rot, CaS sprays gave better control of peach and apple diseases than lime-S and Bordeaux sprays, and did not burn the fruit or foliage even under adverse weather conditions. K D JACOB

Influence of Bordeaux mixture on transpiration. W H MARTIN AND E S CLARK New Jersey Agr Expt Sta, Fifth Ann Report 249-55 (1923), Rev Applied Mycol 10, 46—An increase in transpiration rates in potatoes was noted after application of 5-5-50 Bordeaux mixture. Increased water loss occurred during the night following spraying of potatoes with Bordeaux mixture in soils previously adjusted to 15, 30 and 60% water content. The loss was particularly noticeable in the soils of high water content. Increased losses were observed in the 50 and 30 but not in the 15% moisture series during the day. O E SHEPPARD

Bunt of wheat in western Canada. W F HENRY AND W POPP Sci Agr 11, 201-7 (1930)—The CaH_2O treatment of seed even when badly infested with the *Tilletia* spores gives satisfactory control of this disease. One lb of com CaH_2O is used per bu of seed. $CuSO_4$ and $CuCO_3$ are effective only when the seed is but slightly contaminated with spores. C R FELLERS

Sulfur dusting for the prevention of a bacterial disease of wheat called black chaff. F J GREANEY Sci Agr 11, 274-80 (1931), cf C A 24, 5925—The black chaff disease caused by *B. translucens* was largely prevented or controlled by frequent applications of S dust. The quality of the grain was also improved. C R FELLERS

The influence of sulfur dusting on rubber production. C C AMBENT De Bergcultures 4, 823-5 (1930), Rev Applied Mycol 10, 55—Dusting with S against mildew (*Oidium keteae*) on 2 out of 4 Hevea rubber estates in the Malang district of Java resulted in an increased yield of 22 and 14% over periods of 11 and 12 months, resp. The less productive parts of the estates were selected as controls in the expts., so the results should be accepted with reservations. OREN F SHEPPARD

Sulfuring citrus trees. W A ROUSE J Dept Agr Victoria 28, 732-3 (1930)—Dusting citrus trees with powd S at the rate of approx 2 lb per tree did not consistently increase the yield of fruit. K D JACOB

Defoliation of gooseberries by sulfur-containing sprays. HUBERT MARTIN, J South Eastern Agr Coll., Wye, Kent No 27, 183-5 (1930)—The Leveller variety of gooseberry is susceptible to defoliation when sprayed with either 1 in 60 lime-S, 3% dry mix S-lime or 0.4% colloidal S contg 0.6% of soft soap. Addn of $Al_2(SO_4)_3$ to the lime-S did not diminish the amt of leaf fall. The expts indicated that defoliation is caused by the action of elemental S and is not due to the presence of sol sulfides. K D JACOB

Nicotine in paint for woolly aphid control. LEAON CHILDS J Econ Entomol 23, 883 (1930)—The woolly aphid attacks the callus tissue of pruning wounds, predisposing it to perennial canker infection. It was found that nicotine sulfate added to tanglefoot and tree paint which are applied to the wounds acted as a repellent to the aphids and as a contact poison to the young aphids when they attempted to establish themselves on the callus tissue. C H RICHARDSON

Neonicotines and certain other derivatives of the bipyridyls as insecticides. C R SMITH C H RICHARDSON AND H H SHEPPARD J Econ Entomol 23, 863-7 (1930)—Twenty five bipyridyl derivs and related compds not previously reported have been prep and examd as contact insecticides. These include a no of isomeric bipyridyls, bipyridyls and pyridyl piperidines. Neonicotine (β pyridyl α -piperidine) was the most toxic of these compds, comparing closely with nicotine, to which it is chemically similar. α Pyridyl β piperidine stands next in toxicity to neonicotine of the compds investigated. In general, the compds with the α β and β , α groupings lead in toxicity over compds with rings located in other positions. C H RICHARDSON

Results of airplane dusting in the control of cotton boll worm (*Heliothis obsoleta*

Feb.). FRANKLIN SHERMAN *J Econ Entomol* 23, 810-3 (1930) —Large-scale airplane dusting operations with Ca arsenate in Texas showed that the cotton boll weevil might be controlled by the use of 5-6 lbs Ca arsenate per acre. Cotton boll worms, however, were not controlled by this treatment, but increased in the dusted area during the season. Cotton boll worm damage was greater in treated than in untreated areas. C H R.

Laboratory tests of miscellaneous chemicals against the codling moth. L. C. McALLISTER AND E R VAN LEEUWEN *J Econ Entomol* 23, 907-22 (1930) —This is a study of the toxicity of a large no. of compds to the newly hatched larvae of the codling moth (*Carpocapsa pomonella*). Although no compd superior in practice to Pb arsenate was found, about 50 were sufficiently toxic to be selected for further trial. Among these acetacetanilide, diazaminobenzene, dibromonaphthalene, 2,4-dinitrophenol, 3,5-dinitro-*o*-cresol, nitronaphthylamine, veratrine and nicotine were highly toxic to the larvae. C H RICHARDSON

Further results with trap baits for capturing the codling moth. M A YOTHERS. *J. Econ Entomol* 23, 923-9 (1930), cf *C A* 24, 4351 —Tests with trap baits to capture adults of *Carpocapsa pomonella* were made in a heavily infested apple orchard. The most promising baits were malt sirup, cane molasses, beet molasses, brown sugar and geraniol. Variations in temp affected the no. of moths captured more than did variations in diln of the bait. Geraniol increased the attractiveness of brown sugar and beet molasses baits. The capture of moths per trap increased with the no. of trees available from which to attract them, but not in direct ratio. C H RICHARDSON

Control of weeds by sodium and calcium chlorates. J W DEEM *New Zealand J. Agr* 41, 1-3 (1930) — NaClO_3 seems to give better control of weeds than $\text{Ca}(\text{ClO}_3)_2$, and it is cheaper and easier to handle. $\text{Ca}(\text{ClO}_3)_2$ readily absorbs moisture from the atm., but its fire risk is less than that of NaClO_3 . Most soft weeds are killed by one application of either of these materials, while the harder weeds such as California thistle, blackberry, etc., are greatly weakened. K D JACOB

Eradicating perennial weeds with chlorates. A C ARMY, R O BRIDGFOED AND R. S. DUNHAM. Univ Minn, Agr Expt. Div., *Circ* 32, 4 pp (1930), cf *C A* 24, 2538. C R FILLERS

Rotenone as a contact insecticide. W M DAVIDSON *J Econ Entomol* 23, 868-74 (1930). —Rotenone, the most toxic ingredient in *Derris* and some other plants, was tested in aq suspension and in dust form (mixed with diatomaceous earth) against a no. of species of insects. The aq suspensions were highly toxic to aphids, thrips, white fly larvae, leaf-hoppers, larvae of beetles, tent caterpillars, etc. Adult beetles, squash bugs, red spiders and mealy bugs were more resistant. The dusts were effective against chicken lice, roaches and cabbage worms, but the results obtained against soft-bodied sucking insects were not so good. C H RICHARDSON

The relative value as contact insecticides of some constituents of *Derris*. W. M. DAVIDSON. *J Econ Entomol* 23, 877-9 (1930) —Aq suspensions of the 4 principal toxic constituents of *Derris* root, i. e., rotenone, deguelin, tephrosin and toxicolol, were tested as contact insecticides against aphids, thrips, white fly larvae and red spider mites. Their relative toxic values to *Aphis rumicis* stand with rotenone first in the approx. ratio of 400:40:10:1. Rotenone and deguelin are more toxic than nicotine to *A. rumicis*. This study indicates that the toxicity of *Derris* preps. is due very largely to their rotenone content. C H RICHARDSON

Petroleum insecticides. C. W. WOODWORTH *J. Econ Entomol* 23, 848-51 (1930). —Kerosene emulsion has now been largely replaced by emulsions of the heavier petroleum oils. Heavy oils are selective in toxicity, systemic in their poisonous effects and are absorbed through the tracheae. Among the more than 100 com brands of heavy oils now on the market, control labs generally distinguish the following types: (1) miscible oils made with a cresol soap, (2) soap emulsions made with ordinary soaps; (3) the nonsoap emulsions commonly emulsified with alk caseinate. Three types of oil are distinguished. (1) the kerosenes, about 40° B ϵ , constituting about 2% of the trade; (2) summer oils about 30° B ϵ , (3) crude and winter oils, about 20° B ϵ . The crude oils consist of a mixt. of oils varying widely in sp. gr. The winter oils are distillates from which the lighter and heavier fractions have been removed; some of them approach the summer oils in sp. gr. and refinement. Manufacturers often classify summer oils as heavy, medium and light, terms which are misnomers, as oils called heavy and light may have precisely the same sp. gr. The terms *thick* and *thin* are proposed for heavy and light. The following specifications for oils are given (the values denote, resp., sp. gr in degrees B ϵ ., Saybolt viscosity, percentage unsulfonatable residue, percentage distd. at 350° and 325°F. and, except in the first case, the percentage evapd. after 2 and 24 hrs.). winter oil, 22° 110 sec., 68, 40, 15%; thick oil, 29° 110 sec., 97, 30, 10, 20,

55%, medium oil, 30° 80 sec, 91, 50, 29, 30, 65%, thin oil, 30° 60 sec, 93, 75, 25, 40, 55%. It is conceivable that these phys characteristics have no real significance in detg the oils' efficiency as insecticides. The asphalt base oils of California, which vary considerably in phys characteristics, exhibit remarkable insecticidal efficiency. A tentative classification of the effect of oils on plants is outlined. Two new terms are proposed: *biolyte*, a general term for insecticides, fungicides etc., and *biolysis*, the killing of pests. *Biolytology* is the science of economic poisons. C H RICHARDSON

Some physical properties of certain dormant oil emulsion-sulfur combinations. M D IARRAR AND M A SMITH. *J Econ Entomol* 23, 979-85(1930).—A petroleum oil emulsion for use in a dormant tree spray is described. This emulsion mixes readily with flotation S to give an effective combination spray for San José scale (*Aspidiotus perniciosus*) and peach leaf curl. The emulsion contains a vegetable gum emulsifier. When the emulsion and the S are mixed directly, the phys properties are superior to the mixt in which excessive water is added at time of mixing. Oils emulsified with a gum in combination with S were superior physically to those emulsions with K fish oil soap, K caseinate and petroleum soaps. Size of the S particles is a minor factor in these mixts. Flotation S contains small quantities of electrolytes which exert an important influence upon the oil emulsion mixt. The most desirable phys properties of the mixt were obtained when 4-6 lb of flotation S was mixed with 1½ gals of the oil gum emulsion and diluted in 50 gals of water. C H RICHARDSON

A note on the relation between insecticidal action and the physical properties of soap solutions. P A VAN DER MEULEN. *J Econ Entomol* 23, 1011-2(1930).—This study is an effort to correlate certain phys properties of soaps with their toxicity to insects. The adult Japanese beetle (*Popillia japonica*) was used as the test insect. As wetting and penetration are greatly influenced by the surface tension of the soln, this property was detd with a Du Nouy surface tension app. Surface tension was found to be very low and in all cases no definite relation existed between the surface tension of the soln and its toxicity. The extent of penetration of a soap soln into the spiracles and tracheae of the beetle was not correlated with its toxicity. There was also no correlation between the viscosity of a soln and its toxicity. When the soaps were allowed to evaporate in the air on a glass surface, the surface became covered in most cases with a film of soap which differed greatly in strength with the different soaps. With the K soap made from palm oil, which was of relatively high toxicity, a very tough film remained upon the glass. The results show a relation between the toxicity to the insect and the tenacity of a film left after evaporation. Soaps which form the toughest and most adherent films are most highly toxic, those which form less tenacious and weaker films are less toxic, while those which leave no film at all do not kill the insects. C H RICHARDSON

Densities of mixtures of air and various fumigants. R C ROAKE AND O A NELSON. *J Econ Entomol* 23, 945-7(1930).—The suitability of a compd for use as a gaseous insecticide depends upon its toxicity, its vapor pressure and its vapor density. The vapor densities (in mass of 1000 cu ft of satd mixt of air and gas at 760 mm Hg and 25°) and sp gr (air = 1) at 760 mm and 25° of 29 compds, which have been used as fumigants for insects, are given in tabular form. Formulas for computing vapor density are also given. C H RICHARDSON

The calibration of flow meters for the measurement of insecticide gases. LYMAN C CRAIG AND C H RICHARDSON. *J Econ Entomol* 23, 988-91(1930).—The calibration of resistance-tube flow meters for both large and small flows of air is described and the app is illustrated. C H RICHARDSON

An additional statement concerning the tank-mixture method of using oil spray. RALPH H SMITH. *J Econ Entomol* 23, 1009-11(1930).—Criticism of S's previous paper (C A 24, 3594) has occasioned these remarks. The tank mixt method probably has an important place in the spraying of citrus trees in S Calif whether it will be applicable to the spraying of deciduous fruit trees remains to be proved. Mech mixts of oil and water were used prior to 1912, but modern spray equipment and recent knowledge concerning the functioning of oil sprays indicate that a return to the original method is justified. The tank mixt method is based upon the use of an oil of given specification as to viscosity, distn range and sulfonation. Insecticidal efficiency and injurious effect upon the plant depend entirely upon the amt of oil deposited during spraying. The amt deposited can be governed by placing the emulsifying material directly in the water in the spray tank, adding the oil and maintaining a uniform mixt by means of agitators. The stability of the emulsion and size of oil globules, factors which heretofore have been the focus of oil-spray investigators become relatively negligible. The dependability of the tank mixt hinges upon the proper agitation

and proper kind of emulsifier. The term 'emulsifier' loses its significance when used in connection with the tank mixt. method because an emulsion in the ordinary sense is not produced. The term 'sprayer' is perhaps more suitable. An investigation of sprayers for tank mixt. sprays has shown that powdered blood albumin is one of the most promising substances for this purpose. C. H. RICHARDSON.

The efficiency of the air-blast type of sprayer for applying insecticides. OLIVER L. SAFF AND JAMES R. THOMSON. *J. Econ. Ent.* 1: 23, 881 (1930).—The relative efficiency of com. air blast sprayers and power sprayer outfits in applying lubricating oil emulsions and lime sulfur sol'n to deciduous fruit trees infested with San José scale (*Aspidiotus perniciosus*) was studied. No significant difference was found in the percent control with lubricating oil emulsion applied with the 2 types of sprayers. With lime sulfur sol'n the power sprayer gave greater efficiency than the air blast type. C. H. RICHARDSON.

Weathering of shell limestone and soil formation near Jena (Horr) 8. Decolorizing reducing fertilizer from sewage, etc. (Brit. pat. 335,682) 14. SeSe_2 [fungicide or insecticide] (U. S. pat. 1,912,738) 15.

Ergebnisse der Agrarkulturchemie, Band II, 1930. Edited by F. HONCAMP. Berlin Verlag Chemie, G. m. b. H. 190 pp. M. 13 bound M. 14.

GRUNTER, J. Die Chemie und das Praktikum für den Landwirt. Leitfaden für den Unterricht Würzburg. Bonitas Bauer. 135 pp. Lanen. M. 3.80.

Mededeelingen van de Landbouwhoogeschool te Wageningen. DL34. Verhandeling. 7. I. Vergleichende mikroskopische, physikalische und chemische Untersuchungen von einem Kalkstein- und einem Löss-Bodenprofil aus den Niederlanden. II. Vergleichendes Studium von einem Kalksteinbodenprofil aus Java. By A. van WICHEM, L. MÖSER AND C. VAN AGGELAN. Edited by J. van Baren. Wageningen: H. Veenman & Zonen. 103 pp. Fl. 4.

Fertilizers. I. G. JARREND AND A. G. BRL. 355,475, June 15, 1929. Various granular fertilizers are prevented from caking by coating the granules with mineral oil or other oil insol. in water and non volatile or but lightly volatile. Materials which may be thus treated include urea, KCl and NH_4NO_3 or their mixts. with NH_3 , phosphate, NH_3 , sulfate nitrate, mixts. of NH_4NO_3 with CaCO_3 and other mixed fertilizers contg. NH_3 , K or Na nitrates.

Fertilizers. LOUIS LONGCHAMON. Fr. 693,306, July 10, 1929. The fertilizing value of phosphorous slags is increased and the grinding thereof is facilitated by dropping the molten slag into water.

Fertilizers. GUNTER LOKTONS. Fr. 693,328, April 8, 1930. A fertilizing, stimulating and parasiticide mixt. to be applied to parts of plants above the ground is composed of Fe 25, Cu 5, Mn (MnO_2) 2, K (K_2O) 1, Mg (MgO) 3, Ca (CaO) 20, P (P_2O_5) 1, S 8 and inert substances and water 27½.

Fertilizers. SOC. D'ETIENNES POUR LA FABRICATION ET L'EMPLOI DES ENGRAIS CHIMIQUES (Georges Chindron and Henri Herlemont, inventors). Fr. 693,319, July 12, 1929. Natural phosphates are made soluble by treatment with HCl or HNO_3 and sep'd from CaCl_2 and $\text{Ca}(\text{NO}_3)_2$ by conversion into insol. compds. such as phosphates of Cu, Sn or Fe. The insol. compds. are afterward converted into sol. phosphates for use as fertilizers. Cf. C. A. 24, 3077.

Fertilizer. A. R. LAMB. Brit. 355,600, June 27, 1929. See Fr. 678,693 (C. A. 24, 3500).

Fertilizer. EDWARD W. HARVEY (to The Harrett Co.). Can. 308,507, Feb. 10, 1931. A fertilizer is produced by treatment of $\text{Ca}(\text{H}_2\text{PO}_4)_2$ with atomized NH_4OH soln.

Fertilizers. J. VAN DER PLANK. Belg. 371,268, July 31, 1930. Raw manganiferous Ca phosphates are roasted in an oxidizing atm. Cf. C. A. 24, 4888.

Basic phosphate slags. HENRY AND SCHULWERK HENSEN A. G. and FRIEDRICH HEINRICH. Fr. 692,510, Mar. 21, 1930. A compost of basic phosphate slags is made (1) by decanting Thom's flour or other basic phosphate slags in water, adding an acid or acid soln. and evapg., (2) by wet grinding in the presence of acids or acid salt solns., (3) by disaggregation with acids under high pressure or (4) by the action of finely pulverized acids or acid salt solns. on the slags. Fr. 692,511 describes the selective elimination of CaO from basic slags by treatment with acids or acid salt solns. so as to leave unmodified the constituent called subocarnotite which is the solv. vehicle in citric acid.

Spray oils. WALLACE J. YATES (to Shell Development Co.) Can 208,745, Feb. 17, 1931 A spraying liquid for insecticidal purposes consists of an emulsion in water of a petroleum spray oil, to which not over 1% of S has been added. Cf. C. A. 24, 5923
Insecticide. STANDARD OIL DEVELOPMENT Co. Fr 602,568, Mar. 2, 1930 An insecticide is composed of a naphtha or petroleum distillate having a b p 200°-205°, preferably between 215° and 230°, to which may be added exts. of pyrethrum, etc. Cf. C. A. 25, 1323

Insecticide and fumigant containing ethylene oxide. RURIC C. ROARK and RICHARD T. COTTON (to the Government and People of the U S A) U. S. 1,791,429, Feb. 3 A compn. suitable for fumigating foods, etc., comprises ethylene oxide as essential active agent.

Fungicides and bactericide. JOHN W. ROBERTS. U. S. 1,791,430, Feb. 3. A compn. suitable for combating fungi causing peach scab, etc., comprises a mixt. formed from $ZnSO_4$, 4 lbs., lime 3-4 lbs. and water 50 gals. Casein or album may be added

Fungicides for seeds. I. G. FARBENFAB. A.-G. (Wilhelm Schepss, Wilhelm Bonrath and Carl Taube, inventors) Ger. 515,075, Sept. 23, 1927 Compds in which Hg is linked to an aromatic hydrocarbon radical are used alone or with other fungicides, wetting agents or diluents Among the compds specified are $PhHgOAc$ and CuH_2HgCl . Cf. C. A. 24, 2539

Fungicides for seeds. WILLI LEVY. Ger. 515,098, June 7, 1925 Water-insol Cu compds are mixed with alkali carbonates or bicarbonates and with hygroscopic compds. Adhesives, diluents, etc., may be included, but water-sol fungicides should not be present. A suitable compn. contains $Cu(OH)_2$, 170, $CaCl_2$, 30, $NaHCO_3$, 50 and talc 50 parts.

Immunizing plants. SOC. D'ÉTUDES SCIENTIFIQUES. Fr. 693,723, July 18, 1929 Plants are immunized against the action of bacteria by the use of vaccines or artificial enzymes. Vegetable scrums may be made by vaccinating plants and extg and may be used in human and veterinary therapeutics

Immunizing seed grain and preserving wood, glue, etc., from decay. I. G. FARBENFAB. A. G. Brit. 335,527, May 24, 1929 Fungicidal org. Hg compds are used such as methoxyethylmercuric chloride or acetate, ethoxyethylmercuric chloride, methoxypropylmercuric acetate, phenoxyethylmercuric chloride, benzyloxyethylmercuric chloride, methoxyethylmercuric oxalate and methoxyethylmercuric hydroxide (details for the prepn. of some of which are given) and these compds may be mixed with dry pulverulent diluents or with liquid vehicles Cf. C. A. 24, 2828

Mordants for seeds and for preserving and disinfecting grains, etc. I. G. FARBENFAB. A. G. Fr. 693,415, Apr. 5, 1930 Use is made of Hg compds of the formula A (A) $XHgCRRRCRRR'$, (B) $XHgCRRRCRR-$, in which X is a hydroxyl or any group forming with Hg a salt or a complex salt (acetate, lactate, oxalate, sulfate, chloride, thiocyanate, etc.) R is H, alkyl, aralkyl or aryl and may be the same or different in each position, and R' is an alkyl, aryl or the group B

Herbicides. I. G. FARBENFAB. A.-G. Fr. 692,996, Mar. 28, 1930 Weeds, etc., are destroyed by solns. of salts such as chlorates to which are added a soln. of a wetting agent such as an aromatic sulfonic acid, the amt. of wetting agent added being small relative to the amt. of herbicide

Weed killer. CHEM. FAB. LUDWIG MEYER. Ger. 515,414, Feb. 6, 1929 Addn. to 441,213 and 478,446 (C. A. 23, 4295) Solns. of heavy metal nitrates contg free NO are used. Thus, 0.1-0.5% of NO in the form of HNO_2 or fuming HNO_3 may be added to a 1.5-7% soln. of $Cu(NO_3)_2$

Destroying beet worms. CHEM. FAB. VON HEYDEN A.-G. (Walther Baunacke, inventor) Ger. 515,346 and 515,347, July 10, 1928 Formic acid is used, e. g., by applying to the soil a mixt. of $HCOONa$ 1 and $KHSO_4$ 2 parts (515,346) Picric acid may also be used and may be applied to the soil mixed with diluents, fertilizers, etc., or liberated in the soil by using mixts. of picrates with acid compds (515,347).

16—THE FERMENTATION INDUSTRIES

C. H. FREY

How much alcohol is produced by yeast? II. Checking of alcohol production. STAGER AND GLACHTIZ. *Brenner-Ztg* 47, 127(1930); cf. C. A. 24, 4890—Ale was added to the original mash in various amts. (6, 8, 10 and 14%) and sugar was then added (75, 50, 25 and 0 g) The alc. produced after 21 days was detd The results were

10.6, 8.2, 6.0 and 2.2%, showing the the alc.-producing ability of the yeast is checked when a certain concn is attained. III. *Ibid* 116-7—Various amts. (25, 30, 35 and 40 g.) of distillery, beer and wine yeast were used as seed. The results showed that the alc. production did not increase in spite of the increasing amts. of seed yeast, although the fermentation was completed sooner. S. JOZSA

Alcohol yields from corn and durrs and from damaged raw material. STAIGER. *Brenner-Zig* 47, 114-5 (1930). Ten corn samples, 2 durra samples and 12 miscellaneous grain samples were analyzed and the results of yield of alc. are given. S. J.

The balance in the fourth form of fermentation in the cell-free yeast fermentation. MARIA KOBEL and MAX SCHULZ. *Biochem Z* 229, 238-47 (1930).—In expts in which still yeast maceration juice acted upon hexosediphosphate pyruvic acid was isolated which was equiv. to 70% of the carbohydrate used and actually 100% of the sugar transformed. S. MOSGULLS

Further studies on the formation of methylglyoxal and pyruvic acid by yeast under the influence of various plasmolytic substances. CARL NEUNBERG and MARIA KONRL. *Biochem Z* 229, 275-62 (1930). cf. C. A. 25, 123.—The formation of CH_3COCHO from hexosediphosphate can be easily demonstrated in the presence of various plasmolytic agents (toluene, bromobenzene, various alcs., CHCl_3 , CCl_4 and urethan) or of a high concn. of the diphosphate itself. The accumulation of CH_3COCHO can also be brought about by these means. S. MOSGULLS

Decomposition of non-phosphorated sugar by yeast with the formation of glycerol and pyruvic acid. CARL NEUNBERG and MARIA KONRL. *Biochem. Z* 229, 410-51 (1930).—The fermentation of the Mg salt of fructose, diphosphate as substrate, yields 100% of the theoretical amt. of glycerol and pyruvic acid. Both these substances are recovered, though in smaller quantities, even when a 10% glucose soln. contg. 0.125-0.75% $\text{Mg}(\text{PO}_3)_2$ is fermented by fresh press yeast. The $\text{Mg}(\text{PO}_3)_2$ can be replaced by NaH_2PO_4 . However, since MgO acts just as well, the presence of the phosphate ion is not essential, nor is the Mg ion indispensable because the Mg salt can be replaced by NaH_2PO_4 . The glycerol and pyruvic acid appear in equiv. amts. under conditions where no cell multiplication takes place (yeast juice or suspension preserved with anti-septics), but the equivalence disappears when one works with living yeast cells. Cf. C. A. 24, 1308, 3078. S. MOSGULLS

Formation and identification of acids produced by different strains of propionic acid bacteria. P. W. WILSON, E. R. FIRD and W. H. PATERSON. *Biochem Z* 229, 271-80 (1930).—17 strains of propionic acid bacteria were examd. from the standpoint of their ability to ferment glucose, lactose and maltose, from which volatile acids were produced. In pure cultures sugar was used up only partially but in the presence of *L. casei* all the sugar was used up and 25-75% converted to volatile acids. The most active strains were employed in the fermentation of meal from malted corn, molasses or hydrolyzed starch, only the latter 2 being quickly fermented. The activity of the different strains varied, of course, but those fermenting pure sugar most effectively did so also with the natural sources of sugar. Yeast water or the residue in the distn. of acetone BuOH fermentation served well as the N source. Acetic and propionic acids were isolated and identified as the volatile products, the latter constituting 65-70% of the total. S. MOSGULLS

Studies on slops. STAIGER. *Brenner-Zig* 47, 142 (1930).—The moisture, acidity and alc. content were detd. on 9 different samples of residual liquid from distn. of alc. liquors. The protein, fat, crude fiber, N-free ext. and ash were detd. on 3 rye slops. S. JOZSA

Slop and its significance in agriculture. K. G. SCHULZ. *Brenner-Zig* 47, 162 (1930). S. JOZSA

Removing wood taste from wine brandies. C. LUCKOW. *Brenner-Zig* 47, 139 (1930).—Two samples of wine distillate were stored in wooden barrels and had a strong oak-wood taste. By using 10 g. gelatin per 100 l. distillate, or 1% skim milk, the undesirable taste was practically removed. S. JOZSA

Formic acid, a constituent of the volatile acids of wine, and its determination. W. SHIPST and MASIE HENSCH. *Expt. Sta. Klosterneuburg, Denkschrift*, 70th anniversary, 148-63 (1930), cf. Krepz, C. A. 5, 1305; Placke, C. A. 5, 1300, 3705; 6, 1787; 7, 2258.—The method of F. for detg. HCOOH is very exact, but laborious; the KMnO_4 method of Lieben (1893) was here adopted, and the results of the examn. of 21 samples of Austrian and Hungarian wines are tabulated. Conclusions. (1) The content of free HCOOH varies from 0.023 to 0.080 g. per l.; the quantity bound as ester is greater or smaller; no regularity is apparent. (2) The ratio of free HCOOH to volatile acids shows very wide limits; 3.4-21.0% in grape wines; 5.2-9.5% in raisin wines; and 7.1-

10.5% in fermented fruit juices. (3) The highest content of HCOOH is found in 1 of 2 Tokay wines examd. 0.089 g free, 0.214 g total per l. K records for Tokay wines 0.173–0.517 g total HCOOH per l. These wines are usually left on the lees for about 1 year, then are stored in incompletely filled barrels in order that the characteristic "bread" taste may form. Yeast activity is thus renewed and volatile acids are formed. In the decompn. of leucine of yeast, HCOOH is split off (P. Ehrlich, 1905). (4) Two raisin wines prep'd by S and U contained only 0.108 and 0.091 g per l of total HCOOH probably because 4 g tartaric acid per l had been added to exclude secondary fermentations. Khoudabachian (1883) found in a raisin wine 0.94 g AcOH and 0.28 g HCOOH per l, probably caused by the presence of bacteria. (5) The undiluted juices of currant and raspberries fermented with pure yeast contained when examd. fresh only 0.052 and 0.062 g total HCOOH , whereas K had observed in stored raspberry wine 0.323–0.542 g per l. The cause of this difference will be further studied. Fruit juices apparently keep best when they are fermented with pure yeast, withdrawn from the sediment early and filtered. S WALDBOTT

Determination of glycerol in wines and fermented beverages. LUCIEN SEMICHON and MICHEL FLANZY. *Ann. fals.* 23, 583–602 (1930); cf C A 24, 2231.—S and P recommend the following technique, which is a modification of that of Ferré and Bourges (C A 22, 4713). Place 20 cc of wine in a 100-cc. volumetric flask, make up to 100 cc with neutral 95% alc., add to 5 g BaO and triturate in a mortar for 2–3 min., return to the volumetric flask, let stand 30 min and filter (preferably through a Gooch crucible). Pass CO_2 for 2 min. through a 50-cc. aliquot in a 100-cc. Erlenmeyer flask, evap. to about 10 cc., add 20 cc. water, evap. to 15–20 cc., let cool, make up to 50 cc and filter. Steam-distill 25 cc. (≈ 5 cc. of original sample) under a partial vacuum of 20–40 cm. of Hg, absorbing the distillate in a mixt. of 10 cc CrO_3 (63 g per l) and 20 cc of 60% $\text{Be H}_2\text{SO}_4$. Immerse the tube contg. the sample to be distd. in a vaseline-oil bath heated to 115–20°, which also contains a coil for superheating the steam used for the distn., place the absorbing soln. in a 500-cc. Durand washing flask immersed in a boiling water bath. After distg. for 3 hrs., titrate the excess CrO_3 with $\text{Fe}(\text{NH}_4)_2(\text{SO}_4)_2$ soln. (142.13 g per l), using $\text{K}_3\text{Fe}(\text{CN})_6$ as outside indicator. 69 mg glycerol is completely oxidized by 350 mg CrO_3 . Details are given of the investigation of the method, and data are furnished justifying the various steps of the technique. A. PAPIEAL COUTURE

Experiments toward determination of sorbitol according to Werder and Zsch. RUDOLF HAD and FRANZ POPPEGER. Expt. Sta. Klosterneuburg, *Denkschrift*, 70th anniversary, 90–3 (1930), cf Werder, Zach and others, C A 23, 928, 4296, 24, 5105.—Heretofore, not much less than a 10% admixt. of fruit wine to grape wine could be detected by isolating dibenzalsorbitol (A). A 3% addn. contg. 7.5–39 mg sorbitol per 100 cc. (cf Fellenberg) should yield about 15–73 mg of A, of which only 5–10 mg is needed for the Tutin identification test. However, a 3% addn. yields only 3–5 mg, water insol. A. To obtain a max. yield of A, the temp. must be kept at 0°, a min. concn. of 1 l H_2SO_4 and a certain excess of BzH must be used, e.g., 30 mg sorbitol, 0.65 cc H_2SO_4 (1.1) and 0.2 cc. BzH (6 drops) at 0° yielded 54 mg of A. Slight deviations alter the yield considerably. S WALDBOTT

Brewing as a branch of science. ARTHUR R. LING. *Brewers J* 65, 302–6 (1930).—The biochem. processes involved in the brewing of beer are outlined. P J F W

English barleys of 1929. H M CHURCH. *J Inst. Brewing* 36, 221–4 (1930).—In contradistinction to the heavy barley of the 1928 season, those of 1929 are light, the greater bulk reducing the malt house output by 5%. Germination on the floor was good and regular. Diastatic power and cold water exts. of the malts were below normal. PETER J F WEBER

Foreign barleys of 1929. T R. SUTCLIFFE. *J Inst. Brewing* 36, 225–6 (1930).—California barley malted well and gave high to abnormally high exts., while Chile produced a barley of av. quality. The crops of other barley-growing countries were of inferior quality. PETER J F WEBER

Some recent advances in the chemistry of enzymes. R. H. HOPKINS. *J Inst. Brewing* 36, 189–94 (1930).—See C A 24, 2764 (the vol. no. given for *Chem. News* is an error for 2443). PETER J F WEBER

Cleaning and disinfecting agents (in the brewery). P. PETIT. *Brasserie & malterie* 20, 337–42 (1930).—A brief discussion of the proper method of cleaning and disinfecting piping and bottles. A. PAPIEAL COUTURE

Attenuation. J DE CLERCQ. *Bull. assoc. anciens élèves de Louvain*, March, 1930, *Brewers J*, 66, 293–4 (1930).—The factors influencing attenuation of fermenting liquids are discussed. PETER J F WEBER

The brewing value of hop tannin. A. A. D. COMBIE. *J Inst. Brewing* 36, 307–11

(1930) — Research thus far indicates that boiling a malt wort renders part of the nitrogenous constituents insol. The tannin of the hops accelerates pptn through the formation of tannin nitrogen compds, which however only become quite insol by conversion of the tannin into its deriv., phlobaphene. The tannin nitrogen compds sometimes pass into the beer and can produce haze of two kinds: reversible (when the haze is reduced by chilling) or irreversible (when the tannin is converted into the insol. phlobaphene by heating, such as pasteurization). PETER J. F. WENER

The solubility of maize proteins in mashing. II. R. H. HOPKINS. *J. Inst. Brewing* 36, 297-304 (1930) cf. *C. A.* 24, 1929 — Lab. fermentations of malt and malt maize worts having similar ds and contg. approx. 50 mg. of assimilable N per 100 cc. justified the conclusion that the N derived from maize is almost wholly left in the beer after a single fermentation. A second seeding of the fermented worts with yeast indicated that the maize N would be ultimately assimilable. PETER J. F. WENER

Respiration and fermentation of top and bottom beer yeast. KURT TRAUTWEIN AND JOSEPH WASSERMANN. *Biochem. Z.* 229, 124-53 (1930) — The rate of respiration of top yeast is on the av. 77% higher than that of the bottom yeast. The top yeasts rise to the surface during fermentation where they can easily find O_2 and thus become acclimated to an aerobic existence. But here the increase in the respiration capacity is accompanied by a similar increase in the fermenting capacity so that the ratio between their rates remains const. Both types of yeast, however, use up 2-3% of the total sugar undergoing dissimulation in the respiration and 98-97% in the fermentation. If the energy value of both the respiratory and fermentative processes is called the total metabolism of the top yeast is found to be 42% higher than that of the bottom yeast. In the top yeast 40% of the total metabolism is respiratory and 60% fermentative, while for the bottom yeast these values are 32 and 68%, resp. S. MOROGLIS

Research work in the yeast field. L. H. LAUFITT. *J. Inst. Brewing* 36, 250-60 (1930) — A review of results obtained, and of problems to be solved in the study of the life and activity of yeast. PETER J. F. WENER

Alcohol losses in yeast manufacturing. F. WAGNER. *Brenner-Ztg.* 47, 162-3 (1930) — The losses on alc. due to aeration were studied on 3 different molasses mashers diluted 8, 16 and 20 times, resp. The amt. of air per cu. m. liquid was 25, 70 and 70 cu. m., resp. The alc. content of the air leaving the system was detd., and it was found that the losses of alc. increased with stronger aeration, and decreased with lower alc. content. S. JÓZSA

Losses of water-soluble phosphoric acid by the clarification of molasses under acid conditions and heat. O. HUMMER. *Brenner-Ztg.* 47, 142-3 (1930) — The loss of water sol. P_2O_5 increases rapidly with decreasing acidity. S. JÓZSA

Korean koji (a kind of the so-called Chinese yeast). HIROSUKE NAGANISHI. Abstract from Rept. Central Lab. S. Manchuria Railway Co. 1929, 41-2 — Thirty seven mold fungi, 9 yeasts and 4 bacteria were found in this prepn. The diastatic powers of the molds are detd. V. F. HARRINGTON

Involution cultures of yeast t. A. T. HENLEY. *J. Inst. Brewing* 36, 304-7 (1930) — To obtain surface growths of yeast which shall be characteristic and help in the differentiation of species, specially prepd. wort-carrageen moss (*Chondrus crispus*) in Erlenmeyer flasks was inoculated with a wild yeast (*S. pastorianus*) and a culture yeast (Dublin No. t). The growths thus obtained were distinctly different from each other, since the colonies were able to grow freely along any individual lines they may have possessed and yet could not sink into the body of the media. Under these conditions a high percentage of mycelial cells was formed and gave rise to the term *involution culture*. The gel produced by the addn. of carrageen moss to wort was fragile, only a little agitation breaking it down. Details of the method of prepn. are given. P. J. F. W.

The use of microorganisms in certain commercial processes (HUBERT) 23. The preservation of sweet must (MEHLITZ) 12.

VENTRE, J.: *Traité de vinification pratique et rationnelle*. Tome I. Le raisin et les vinifications. Paris. Dunod. 490 pp. F. 50

Fermentation products. ZELLSTOFFFABRIK WALDHOF and MAX GADE. Fr. 693,553, April 8, 1930. In the distn. of alc. from a fermented sulfite wort the first running contg. aldehyde is directed to a fresh wort about to be fermented.

Dehydrating alcohol. Wm. H. ENGELS (to Merck & Co.). U. S. 1,790,907, Feb. 3. A hydrous alc. is mixed with comminuted CaO in relatively small excess as proportioned

to the amt of water to be removed, the mixt is heated and refluxed and the film of hydrated lime is continually removed as it forms upon the surface of the particles of oxide by the action of mech agitation, this operation is continued to completion of the dehydration and the treated ale is distd off under continued agitation. A "com pure" hydrated lime is obtained as a by product.

Butanol and acetone. AUGUSTE FERNBACH. Fr 693,744, July 23, 1929. In the production of butanol and acetone by the fermentation of materials contg saccharose, inversion of the saccharose is obtained by the action of sucrase or invertase, yeast being used as the source of sucrase. The medium is kept at a pH value of 4.3-4.5.

Lactic and acetic acids. WISCONSIN ALUMNI RESEARCH FOUNDATION. Brit. 335,506, June 20, 1929. Lactic and acetic acids are produced by fermentation of pentoses and hexoses (such as those from wood, sawdust, straw, corncoals or the like) by the action of a described microorganism obtained from fermenting plant material such as silage or sauerkraut. Various details of procedure are described.

Glycerol. VERRINGE CHEM WERKE A-G. Ger 514,395, June 22, 1926. In the manuf of glycerol by fermenting sugar in alk soln in the presence of Na_2SO_4 , the volatile fermentation products are distd off, and the residual mash contg glycerol is fermented again after addn. of more sugar, these steps being repeated as required. The yeast should preferably be filtered off before removing the volatile products. Examples are given.

Yeast. KRAUSZ MOSKOVITS VERRINGE INDUSTRIE-ANLAGEN A-G. Fr 692,546, Mar 21, 1930. Yeast of high enzymic activity is prepd from yeast multiplied according to any known process, which is further developed in 2 working stages in such a manner that in the one stage, by assimilation, nitrogenous substances complete the plasma of the yeast cells with simultaneous inhibition of multiplication and in the other stage, by setting up a strong fermentation, the enzymes of the cells are caused to develop, the 2 stages being effected in either sequence.

Preparation of air yeast and other microbiological processes employing the passage of gas through the cultura medium. HENDRIX C. JANSEN. Dutch 2,653, Sept 15, 1930. Substances which are to be added to the nutrient medium are mixed in with the gas passed through the medium. Liquids or solids are sprayed into the gas.

17—PHARMACEUTICAL CHEMISTRY

W. O. EMERY

The chemical examination of *Sida cordifolia*, Linn. SUDHAMOY GHOSH AND ASHUTOSH DUTT. J Indian Chem Soc 7, 825-9(1930).—Berela (*S. cordifolia* L.), regarded as a most valuable drug in the Ayurvedic system of medicine, and used as an aphrodisiac by the Mahomedan hakims, has not previously been examd systematically. Assays of the entire plant by the U S P method for belladonna showed the presence of 0.085% total alkaloids, the seeds contg 4 times more alkaloid than the leaves, stems or roots. Extn with petr ether, Et_2O , abs alc and H_2O showed the presence of fatty oil, phytosterols, resins, resin acids, mucins, KNO_3 and alkaloids, but no tannin or glucoside. The alc. ext contained the whole of the alkaloid content. From 100 kg of the air-dried drug a crude alkaloid residue was extd, $\frac{1}{4}$ of which was CHCl_3 -sol. The remainder gave a cryst HCl salt, m 215.5°, α_D^{25} -35.0, chloroplatinate m 190.5°, mol wt 165, mixed m p with ephedrine HCl from *Ephedra vulgaris* 215.5° (ephedrine- HCl , m 215-6°, α_D^{25} -32.5, mol wt 167). The CHCl_3 sol fraction was non-cryst. but appeared to contain alkaloids related to ephedrine. Ppts. with 11 alkaloidal reagents and the violet color of the biuret reaction confirmed the identity of the alkaloid from Berela with ephedrine. Pharmacologically, the alkaloid hydrochloride causes marked and persistent rise of blood pressure in the absence of ergotoxine, dilation of the bronchioles and inhibition of intestinal movements, these effects resemble those of sympathomimetic bases such as adrenaline or ephedrine. C. R. ADORNALL

The distillation of essential oils. V. I. VARENYTSEV. J Chem Ind (Moscow) 6, 1602-4(1929).—The extn. of essential oils from their seeds by steam distn can be expressed by the formula of an infinite decreasing geometrical progression $S = a/(1-r)$ and $t = ar^{n-1}$. Thus, by experimentally establishing a (the amt. of oil extd in the first hr) and r (the ratio of the amt. of oil extd between 2 consecutive hrs) the duration of the distn. of the oil content of the seeds can be easily calcd. The validity of this "distn. law" is shown by numerous examples taken at random from the literature.

E. BIELOUSS

The seeds and oil of *Sinapis dissecta* Lag. N. BILVAPV. *Masloboino Zhiroroe Delo* 1929, No. 6, 25-6. —This plant, found abundantly in the Saratov region, is a variety of the common *Sinapis alba* L. The seeds contain moisture 8.80, oil 28.05, fiber 15.05, ash 4.48, total N 3.63, crude protein 23.90 and essential oil 0.09%. The oil has d 0.9150, solidification point 15, n_D²⁰ 1.4719, sapon no 172.4, I no (Hahl) 100.52 acid no 1.64. The fatty acids have I no 99.38, m 21°, solidify 18°. The oil is thus similar to the oil of *Sinapis alba*. I. BIRLOUSS.

Investigation of the essential oil from different kinds of fennel cultivated in the Krasnodar district. N. SOBANYIN AND S. SAAKOV. *Masloboino Zhiroroe Delo* 1929, No. 6, 34. 6. —Fennel oils of different origins vary in yield from 1.21 to 5.3%, in optical rotation from +5 to +20.3, in α from 1.528 to 1.5405, in solidification point from 3° to 12°. I. BIRLOUSS.

Comparative value of metaphen in alcohol-acetone-aqueous solutions in the pre-operative disinfection of the skin. WYNFIELD W. SCOTT AND KONRAD E. BIRKHANG. *Ann. Surgery* 93, 587-97 (1931). —A 0.5% alc acetone aq soln of 4 nitro 5 hydroxymercuri *o*-cresol is a satisfactory pre-operative skin disinfectant. RACHEL BROWN.

Note on the biological assay of tincture digitalis. G. A. GRANT AND S. G. ALEXANDER. *Proc. Trans. Nova Scotia Inst. Sci.* 17, 211-7 (1930). —The strength of several samples of *T. digitalis* was tested by the bioassay outlined at the Geneva Conference. The results obtained indicate that for the tinctures examined the max deviation from the U. S. P. X standard is much less than that previously found and that the strength of the International powder is considerably above that of the present U. S. P. X standard. RACHEL BROWN.

The germicidal action of ethereal oils. SIEGFRIED L. MALOWAN. *Z. Hyg. Infektionskrankh.* 112, 03-4 (1931). —The chem. compn as well as the H₂O soly affect the germicidal action of ethereal oils. Oils contg aldehyde are most active and have a high soly; oils contg esters are less active than alcs. RACHEL BROWN.

Bulgarian Otto of Rose. ERNEST S. GUENTHER AND ROBERT GARNIER. *Am. Perfumer* 25, 417-20, 479-82, 547-50, 621-4 (1930). —An exhaustive review is given of the natural rose oil industry as practiced in Bulgaria, with illustrations. W. O. E.

Leaf oil from *Dacrydium frankii* Hooker. A. R. PENFOLD AND J. L. SIMONSEN. *J. Proc. Roy. Soc. N. S. Wales* 63, 05-101 (1930). —A renewed study of the hydrocarbon portion of the oil would indicate that the alleged diene previously reported by Baker and Smith is in all probability identical with Δ^4 -carene, and should, therefore, be removed from the literature until evidence of its sep. existence is forthcoming. W. O. E.

Essential oil of *Eucalyptus rariflora*. A. R. PENFOLD, C. B. RADCLIFFE AND W. F. SHORT. *J. Proc. Roy. Soc. N. S. Wales* 64, 101-14 (1930). —Distn. of the leaves and terminal branchlets of this tree gave a reddish yellow to yellowish brown oil in 2.5% yield, possessing a pronounced phellandrene odor modified by the aromatic aldehydes present. The principal constituents so far identified are the terpenes Δ^4 -carene, β -phellandrene, *l*- α -pinene, β -pinene, cymene, with cineole (about 10%), sesquiterpenes (principally aromadendrene), sesquiterpene alcs., with small quantities of the aromatic aldehydes (cuminal, phellandral, cryptal), alkali sol. compds. (unidentified phenols and dehydroangustone (β diketone)). The identification of Δ^4 -carene is noteworthy in that it is the first record of its occurrence in *Eucalyptus* oils. W. O. E.

Oil of amber. T. TUSTING COCKING. *Perfumery Essential Oil Record* 21, 477-8 (1930). —An explt. study is reported of both the true and the com. oils, notably of the consts. obtained in the examn. of the various oil fractions. W. O. E.

Kalk-vitryn. HERMANN ESCHENBRENNER. *Pharm. Presse, Wiss.-prakt. Heft* 1930, 145-6. —A new vitamin-contg. Ca phosphate prepn. is described, in connection with a review of certain chem. tests for the vitamins A, B and D alleged to be present in the above prepn.

Quinine. KARL DOPF. *Pharm. Presse, Wiss.-prakt. Heft* 1930, 167-8. —The history, production and importance of this alkaloid in medicine are discussed. W. O. E.

Examination of pine needle extracts by capillary analysis. H. ESCHENBRENNER. *Pharm. Presse, Wiss.-prakt. Heft* 1930, 177-8. —A discussion is given of the contaminants not infrequently encountered in com. preps. of pine needle exts. in connection with a review of the method specified by the Supplement to the D. A. -B. VI, as also that involving examn. of the capillary pictures obtained under the analytical quartz lamp. W. O. E.

Relation of structure to action of morphine and its derivatives. RUDOLF ZIFFERER. *Pharm. Presse, Wiss.-prakt. Heft* 1931, 1-2. —Attention is directed (by illustration) to

important and valuable derivs of morphine obtained by means of certain relatively simple changes in its mol. structure W. O. L.

Sickness and intoxication through benzene and its derivatives. MAX GÄNEWALD *Pharm. Presse, Wiss. prakt. Hefte* 1931, 25—A discussion is given of the intoxication effects of C₆H₆ and derivs on the animal economy, and of certain precautions to be observed in com. operations with these products W. O. L.

Reaction of phenylum acetylo-sabtylum. H. SZANCER. *Pharm. Zentralhalle* 72, 68-9(1931)—It is shown that this substance with NaNO₂ and concd. H₂SO₄ yields the same color changes as salol W. O. L.

Detection of juglone. ROBERT FISCHER AND FRITZ STALDER. *Pharm. Zentralhalle* 72, 97-100(1931)—In the microsublimation of *Folia Juglandis* via the D. A. B. VI. noyel low crystals result, in fact the official Ger. procedure for the microsublimation of either fresh or stored leaves, or of 8-day old, long shredded pericarps, fails to yield juglone. However, extn. with CHCl₃ of fresh or 3 day old leaves, followed by suitable treatment of the resulting residue, detects the presence of juglone. Similar results are obtained with 8-day old pericarps. Older leaves are free from juglone. α -Hydrojuglone could not be detected in fresh pericarps. It is shown that the yellowish needles reported by Tummann as α -hydrojuglone were in reality juglone. β -Hydrojuglone could not be detected in any of the fruit available W. O. L.

Estimation of *p*-chlorophenol as such and in pharmaceutical preparations. K. FEIST AND F. KLATT. *Pharm. Ztg* 76, 112-3(1931)—The method is based on the procedure of Koppeschaar, involving conversion of the *p*-chlorophenol into chlorodibromophenol (by KBr and 0.1 N KBrO₃ soln. + HCl) followed by treatment with KI, and final titration with 0.1 N Na₂S₂O₃. In mixts. of *p*-chlorophenol and camphor (or menthol) the former is isolated by treatment with NaOH soln. and then titrated as above W. O. L.

Chemistry of neodorm. G. KUHLMANN. *Pharm. Ztg* 76, 113(1931)—This prepn. is α -bromo- α -isopropyl butyramide, it m. 50-1°, resembles menthol and is sol. in the usual org. solvents and in oil. It decomps. if sterilized in aq. soln., but not in oil W. O. L.

Adulteration of oil of cloves with benzyl alcohol. S. KROLL. *Pharm. Ztg* 76, 128(1931)—A sample of oil of cloves possessing apparently normal chem. and phys. properties, as stipulated in the Ger. Pharm., was found on examn. to contain considerable PhCH₂OH as adulterant. The presence of this alc. was demonstrated by the method of Thomas (cf. *Ber. pharm. Ges.* 1, 278(1891)) W. O. L.

United States Pharmacopeial Commission. L. ROSENTHAL. *Pharm. Ztg* 76, 143-4(1931)—An address indicating the salutary manner in which the Commission operates in marked distinction to the rather secretive activities of similar European bodies W. O. L.

Manganous calcium hypochlorite. O. H. KERSCHNER. *Pharm. Ztg* 76, 184-6(1931)—The reddish color noted from time to time in solns. of Ca hypochlorite is shown to result from conversion of traces of Mn into permanganate. W. O. L.

Production of benzonaphthol and *p*-acetamidophenyl salicylate. JULIUS SCHWYZER. *Pharm. Ztg* 76, 186-8(1931)—Detailed procedures are reported for the production of benzonaphthol and salophen in large scale operations, in connection with illustrations of the app. needed. W. O. L.

New method for the preparation of isoeugenol from clove oil. R. PRIESTER. *Rieschstoffind.* 5, 83-5, 108-9(1930)—Eugenol is converted to isoeugenol (1) by heating with KOH, (2) by the action of KOH dissolved in H₂O, MeOH, EtOH or AmOH under normal or increased pressure, (3) by the action of K or Na amylate in AmOH soln. under ordinary or increased pressure. KOH functions purely as a catalyst. According to the new method eugenol or oil of cloves 30 is heated for 35 min. to about 180° and maintained there 1½ min. with H₂O 30, KOH 15, glycerol 5, the mass poured into H₂O, C₆H₆ added, the KOH almost neutralized with dil. H₂SO₄, AcOH added in slight excess, the C₆H₆ soln. withdrawn and washed to neutrality and finally fractionated *in vacuo*. Little or no resinification takes place under the above procedure W. O. L.

Acetoegenol and its presence in clove oil. J. HEROLD. *Rieschstoffind.* 5, 100-2(1930)—The occurrence of acetoegenol in oil of cloves, its probable fate during the process of distillation, in connection with the various methods for evaluation are discussed. W. O. L.

Composite or unitary perfumes. ERNST SCHAFFAN. *Rieschstoffind.* 5, 119-20(1930)—The relative advantages of compns. as compared with unitary chem. substances or essential oils are discussed. The decision must be left largely to the expert knowledge of the perfumer W. O. L.

Chemical and bomb gas baths. W. PEYER AND H. LEFFINGER. *Süddeutsche Apoth. Ztg.* 71, 77 (1931). A general discussion of the essential differences in effect of O and CO₂ bath proper, when developed from salts with suitable catalyzer directly in the water or indirectly from the capillaries of bombs. It is shown that the former method is the preferable one because of the phys. character of the developed gas. W. O. F.

Nature drug plants in the light of modern chemistry and therapy. L. KROENKE. *Süddeutsche Apoth. Ztg.* 71, 93 (1931). cf. C. A. 24, 1931. An address. W. O. F.

Problems of practical pharmacy. HERMANN THOMAS. *Süddeutsche Apoth. Ztg.* 71, 110 (1931), *Pharm. Ztg.* 76, 195 (1931). An address. W. O. F.

American medical plants of commercial importance. A. P. SIVERS. U. S. Dept. Agr., *Misc. Publ.* 77, 1 (1930). About 1500 plants of reputed medical value are described. C. R. FILLERS.

Assay of digitalis. II. Comparison of Mansfeld's sinus method with the six-hour frog or cat method. A. STASIAK AND B. ZBORAY. *Magyar Gyógyszerészeti Társaság Értésítője* 5, 100 (1929). cf. C. A. 23, 1770. Further investigations were made by the method of Mansfeld Horn on spring frogs with powders of digitalis leaves. An infusion made according to Mansfeld was used as standard. Parallel detns. on the same powder agreed well. The differences between the results of the sinus method and those of the 6 hr. frog or cat method are the same as those of the frog method and the cat method between each other. S. S. DE FINÁLY.

Extractum filicis maris. Z. CURKE. *Magyar Gyógyszerészeti Társaság Értésítője* 5, 125-14 (1929), cf. C. A. 23, 4769. No correlation can be found between biol. losses of value and quantity of active matter found by chem. methods. Results of the so-called "pure filicic acid" detns. (Lomme, Kraft) are unreliable, since they do not show the presence of lower piloroglucins which also have anthelmintic power. A modification of the method of Kraft Bloch i romme has been worked out. Shake out the ether soln. of the ext. twice with hot water, and filter. Treat an aliquot part of this soln. with ammoniacal AgNO₃ soln. Let stand for 6-10 hrs. in a dark place. Filter out the reduced metallic silver, wash, dissolve in 20% HNO₃, and titrate with 0.1 N NH₄SCN soln. One cc. of this soln. equals 0.00217 g. piloroglucinol or 0.0196 g. phlorbutyrophrenone. S. S. DE FINÁLY.

Biological evaluation of extracts of Filix mas. Z. CURKE. *Magyar Gyógyszerészeti Társaság Értésítője* 6, 85-101 (1930), cf. preceding abstract. Chem. evaluation alone is not reliable. Biol. tests can be made satisfactorily with *Hirudo medicinalis* and *H. officinalis*. An emulsion of the drug with 1% NaHCO₃ was used. Pots kept closed from the air remain unincorpd. for years. For the test 4 g. of the ether ext. is shaken carefully at 40°, and mixed with powd. guaranine and 1% aq. NaHCO₃ to an emulsion. This is added with 1% NaHCO₃ to 2000 cc. Of this dil. soln. 500 cc. (equal to 1 g. ext.) is again added with NaHCO₃ soln. to 2000 cc. The soln. can be used after 2 hrs. M. I. d. is detd. with three series of expts. One series for orientation, the third for actual detn. M. I. d. found was 0.0028 in 100 cc. The animals should be kept in motion under the detn. by mechanical or elec. means. S. S. DE FINÁLY.

How to distinguish sulfonal from trional. L. EKKERT. *Magyar Gyógyszerészeti Társaság Értésítője* 6, 385 (1930). The substance to be tested is heated with resorcinol and concd. H₂SO₄ until it shows a green color. The liquid is dil. and NH₃ added. Sulfonal shows bright rose color, trional becomes brownish yellow or reddish brown. S. S. DE FINÁLY.

Biological test of adrenaline in draft mixtures, especially in the presence of local anesthetics. A. STASIAK AND L. RIGÓ. *Magyar Gyógyszerészeti Társaság Értésítője* 6, 359-95 (1930). The presence of adrenaline can easily be proved in mixed draughts, even in presence of local anesthetics, by means of surviving mouse uterus. Adrenaline prevents the spontaneous contractions of the surviving uterus. These contractions, however, reappear very quickly after the uterus is washed out with Locke's soln. Thus the same uterus may be used for several detns. By this method adrenaline can be detected in the presence of novocaine, β eucaine, holocaine, perocaine, stovaine and tilocaine and in the presence of ext. of posterior lobe of the hypophysis. Novocaine increases the sensitiveness of surviving mouse uterus to adrenaline, β eucaine and stovaine diminish it. Holocaine showed no influence on the sensitiveness. S. S. DE FINÁLY.

Comparative studies on the tannic acid content of some drugs. I. TEMESVÁRY. *Magyar Gyógyszerészeti Társaság Értésítője* 6, 396-402 (1930). Of the methods studied, (1) international hide powder method, (2) Lowenthal method, (3) SnCl₂ method, (4) Wasieky method, (5) Schulte method and (6) tanniform method, (1) gives the most exact and reliable results. However, this gives high values for cinchona bark (part of

alkaloids is dissolved and pptd by skin powder) and for *Rheum* (where much slime is attached to the hide powder) The tannin method can usually be used when pyrocatechol tannic acid is present
S S DE LVAL

Stability and attenuation of insulin. H PUGGALI, *Rev. soc. argentina biol.* 5, 677-84(1929), *Physiol. Abstracts* 15, 239—Attenuation is avoided by the concn. of the alc. acid exts. of pancreas *in vacuo* at 35-40° Controls were made 200 days later, and no loss of pharmacol. activity was found
G G

New assay method and identification tests for mild and strong silver proteinate. I H INABINETTE, *Am. J. Pharm.* 103, 42(1931)—Tanret's soln. produces a flocculent yellow ppt. when added to an equal quantity of 1% soln. of strong Ag proteinate. With the mild Ag proteinate a flocculent grayish black ppt. is given. Nessler's reagent is even more distinctive when used under the same conditions. A clear, golden brown soln. is produced with the strong Ag proteinate and bluish black soln. with the mild proteinate. Both of these solns. were satisfactory identifying agents for protargentum, protargol, Ag protein, Merck and P. W. R., argyrol, solargentum and argo nuclein (S. K. F. Co.). A new assay method, much more rapid than the present U. S. P. X., is as follows: Weigh 0.5 g. of the Ag compd., dissolve in 20 cc. H₂O, then add 20 cc. concd. HNO₃ and boil for 10 min. Cool, dil. to about 150 cc. and titrate with 0.1 N KCNS, using ferric alum as an indicator. This method produced results identical with those obtained by the U. S. P. X. method.
W. G. GAESSLER

Assay of solution of magnesium citrate. HERMAN W. HAUSSMANN, *Am. J. Pharm.* 103, 44-5(1931)—The present method of the U. S. P. X. for Mg requires evapn. and ignition, a time-consuming operation and one in which error may arise. The following modification of this gives slightly higher but more consistent results. The material after detn. of the free acid is evapd. and ashed as usual. The ash is then transferred to a porcelain dish with hot water and 50 cc. of 0.5 N H₂SO₄ added. It is then boiled, cooled and the excess of acid detd. by titration against 0.5 N NaOH with Me. orange as indicator.
W. G. GAESSLER

Assay method for ointments containing mercury or its salts. R. O. METZLER, *Am. J. Pharm.* 103, 45(1931)—For mercurial ointments the present method as outlined in the U. S. P. X. gives satisfactory and consistent results. 1. or calomel ointment weigh about 1 g. of the ointment directly on a dense filter paper. Fold and pass CCl₄ through until all fat is removed. Place the paper in a glass stoppered flask, add 10 cc. CCl₄, 30 cc. 0.1 N I, 3 g. KI and 10 cc. H₂O. Shake well, allow to stand 12 hrs. and det. excess 0.1 N I with 0.1 N Na₂S₂O₃. Twelve detns. gave results varying from 30.09% to 30.5% calomel. For yellow HgO ointment the following methods were used: (1) Boil with HNO₃, add H₂O and titrate with 0.1 N KCNS. (2) The same as No. 1 except fat is filtered off before titration. (3) Treat as in No. 2 and ppt. Hg as sulfide. (4) Destroy fat with concd. H₂SO₄ and KMnO₄ and det. Hg as sulfide. None of these methods gave results which were consistent or in agreement with the known compn. of the specially prepd. yellow HgO ointment.
W. G. GAESSLER

Determination of boric acid in ointment of boric acid U. S. P. X. THEODORE BUDIN, *Am. J. Pharm.* 103, 46(1931)—Place 25 cc. CHCl₃, 10 cc. glycerol and 1 cc. of phenolphthalein test soln. in a well stoppered container. Shake, add 0.1 N alkali to faint pink color. Then introduce about 1 g. of boric acid ointment weighed on a wax paper, and add 25 cc. of 0.1 N NaOH. Shake well and det. excess of NaOH by titration with 0.1 N H₂SO₄.
W. G. GAESSLER

Investigation of calcium hypochlorite (65 percent available chlorine). BERNARD MELKON, *Am. J. Pharm.* 103, 46(1931)—Analysis of a com. supply of a product marketed under the trade title H. T. H. by the Matheson Alkali Works Inc. gave insol. in H₂O 8.40, SiO₂ 0.28, Fe₂O₃ 0.38, Al₂O₃ 1.76, CaO 32.17, available Cl 62.70%.
W. G. GAESSLER

Assay of chloroform liniment. JOHN C. BRANTLEY, JR., *Am. J. Pharm.* 103, 46-7(1931)—The method of Willgerodt (*C. A.* 20, 92) was compared with a direct standard alkali sapon. and titration of residual alkali. Direct alkali sapon. does not give correct results.
W. G. GAESSLER

Development of pharmaceutical chemistry in Maryland. A. G. DUMEX, *J. Chem. Education* 8, 471-84(1931).
E. H.

The hydroxylamine method for the determination of ketones. Carvone in caraway and dil. oils. C. T. BENNETT and T. TUSTING COCKING, *Analyst* 56, 79-82 (1931)—The method of Bennett and Salamon (*C. A.* 22, 1825) applied to various oils contg. aldehydes and ketones does not give satisfactory results in all cases, the method must be modified according to the particular ketone present. The modified method of the Essential Oil Sub-Committee for the detn. of citral in oil of lemon (*C. A.* 24, 1934) is

of almost universal application for testing oils contg aldehydes, if a little benzene is added as solvent for the non aldehydic portion of the oil but this modified procedure was found to be quite unsuitable for the detn of ketones. After considerable experimentation with dill oil, caraway oil, pennyroyal and pepperstone, the following modification of the method was adopted for detg carvone in oils of dill and caraway. Weigh 1.5 g of oil into each of 2 stoppered tubes, add 10 cc of approx 5% NH_4OH HCl reagent (0.9 g dissolved in 10 cc of 90% alc, 0.4 cc of 0.2% dimethyl yellow added and sufficient N KOH in alc to bring the soln to a dull yellow), and heat in boiling water to about 75°. The color of the soln changes from yellow to red as a result of liberation of HCl. At 5-min intervals remove each tube from the water bath and neutralize with N KOH in alc using a sep buret for each sample. Toward the end, the indicator becomes golden orange instead of red and it is hard to estimate the exact end point. When it is thought that the end of the reaction has been reached add 0.5 cc more of KOH to one of the tubes to overstep the end point and add to the other sufficient NaOH to match the color in the other tube. In the computation, 1 cc of N KOH = 0.1513 g of carvone. The advantages of this method over that of the neutral sulfite method are that very little oil is required and less time. The method has been applied to the detn of ketones in other oils (menthone in oil of peppermint and pulegone in oil of pennyroyal) but the end point was hard to see. W. T. H.

Seventh report of the essential oil sub-committee on uniformity of analytical methods. Determination of solubilities. J. ALLAN *et al* *Analyst* 55, 380 (1930).—The tests should be carried out at 15.5°, unless stated otherwise. The concn of alc should be given in % by vol and the following terms should be used as defined in the report: sol, sol with opalescence, sol with turbidity. Regarding the opalescence and turbidity, the adjectives faintly, slightly and distinctly should be used with reference to expts under stated conditions with AgCl (for opalescence) and with BaSO_4 (for turbidity). W. T. H.

Testing of Admiralty disinfectant fluid. T. C. PATTERSON AND ROBERT C. FRIDRICK. *Analyst* 56, 93-104 (1931).—Two disinfectant fluids were divided into 2 series of identical samples and the phenol coeff by the Rideal Walker method was detd in 13 different labs. The values obtained for one sample ranged from 6.0 to 11.5 and on the other from 6.0 to 17.4. To avoid such discrepancies, the Admiralty method is described in great detail. The method consists in prep a 1% soln of the fluid with artificial sea water, allowing this to stand 21 hrs and then comparing its action in the presence of org matter against a particular strain of *B. typhosus* with similar expts with pure phenol used as reference. A bibliography of 19 titles is appended. W. T. H.

Chlorethane as a preservative. B. POLAK. *Bratislav Bekarske Listy* 9, 231-4 (1929).—Chlorethane ($\text{Cl}_2\text{CClMeOH}$), which is used as a preservative of org substances, is injurious to the circulation by its paralyzing effect on the heart muscle and motor cardiac centers. WILLIAM J. HESA.

Distillation of essential oils. V. G. PATWARDHAN. *Poona Agr. Coll. Mag* 21, 31-3 (1929).—An app for the small scale distn of essential oils from grasses, etc., is described and illustrated. Two to 5 lb of material, depending upon its nature, can be distd at one time. K. D. JACON.

Merthiolate as a germicide. H. M. POWELL AND W. A. JAMIESON. *Am J Hyg* 13, 296-310 (1931).—Merthiolate (I) is $\text{Na ethylmercurithiosalicylate}$, $\text{EtHgSC}_6\text{H}_4\text{CO}_2\text{Na}$, a study was made of the general properties, germicidal and inhibition properties and toxicity for animal tissues. I is a white, orthorhombic cryst compd, readily sol in water and physiol salt soln. Solns of germicidal strength are stable under ordinary conditions of light and temp, are colorless and do not stain tissues or fabrics. The aq solns are miscible with soap and maintain their germicidal property. I does not amalgamate with metals. In soln it is so slightly ionized that there is no immediate pptn of Hg with either alk hydroxides or $(\text{NH}_4)_2\text{S}$. Dry I or aq I can be mixed in various proportions up to 25% with undil serum without causing pptn. I is pptd by the salts of heavy metals and by acids. Many germicidal tests were made and are reported, from these tests it appears to be a very promising germicide with certain unusual properties which make it well adapted to tissue antiseptics. N. A. LANGE.

Estrogenic substances. II. Analysis of plant sources. BURNHAM S. WALKER AND JAMES C. JANNEY. *Endocrinology* 14, 389-92 (1930), cf *C. A.* 24, 3809.—A relatively high concn of estrogenic material occurs in plants associated with green pigmentation. Roots, tubers and fruits are neg. Estrogenic substances in yeast are an exception. Pos extracts are obtained from plants while their growth is most rapid. MARY E. LEAR.

Chemical analysis of the cortex of *Aspidosperma polyneuron* Muell. Arg. Louis

FLORIANI *Ann Farm Biogum* 1, 135-9(1930) —P. made a modification of the procedure of Arata and with its aid found that the cortex of *Aspidosperma polyneuron* Muell Arg contained 2% of alkaloid. The nature of the alkaloid is under study.

B S LEVINE

Studies on the Vitali reaction for the detection of atropine. I. The extent of its applicability. SANTIAGO A CRIST *Ann Farm Biogum* 1, 140-9(1930) —A study of the Vitali reaction as applied to atropine HCl or sulfate, hyoscyamine, scopolamine, isatropylcocaine, homatropine and other alkaloids. The relation between the reaction and the chemical constitution of the reacting substances is discussed. The Vitali and Arnold reactions are compared.

B S LEVINE

Physiological styptics. OSKAR BEYER *Chem-Zig* 54, 1007(1930) —A review of the various styptic agents is made. It is suggested that a rational means of stopping bleeding is to strengthen artificially by catalytic agents, the formation of thrombin. A prepn. of this type would be useful in hemophilia and in patients suffering from jaundice and liver damage. Such a prepn. is made from cow liver. Finely ground fresh cow liver 500 g. is placed in a 2 l flask, one l of 96% alc is added and the contents are extd for 30 min at about 60° on a water bath. The bright yellow liquid is rapidly filtered off on a porcelain suction filter. The filtrate treated with 4 portions of pure ether is allowed to stand 8-12 hrs at room temp. The ether soln is decanted, leaving a yellowish brown mass adhering to the walls of the vessel. The flask is heated to about 30° and the ether removed *in vacuo*. The residue, about 25 g., is taken up in 300 cc sterile water, filtered, and 5 cc put in sterile glass ampules. The ampules, after sealing, are sterilized in steam at about 95° for 1/2 hr. The soln is assayed for clotting power. The substance contains N, peptones, amino acids, purine bases and carbohydrates.

G H W LUCAS

Constitution of amino alcohols as local anesthetics. (Supplement.) S KANAO AND K SHINOZUKA *J Pharm Soc Japan* 50, 1152-62(1930), German abstr 149-52, of C A 24, 3832 —In previous papers K reported on the type I, $RC_2H(OH)CH_2R'$ (NR'R''). The present paper deals with the type II, $RR'C(OH)CH_2NH_2$. The local anesthetic action is neg if R and R' are both Me, Et or Pr, it is pos if R and R' are both Bu, $Me_2CHCH_2CH_2$, Ph or $PhCH_2$, or if R is Bu or cyclohexyl and R' is Ph. The methods of prepn. of these amino alcohols are also given.

F I NAKAMURA

Semen contra pills. J MAHEU AND J CHARTIER *Ann fols* 23, 613-4, *J. pharm chim* [3], 12, 401-3(1930), of C A 21, 2961 —Examn of 4 samples for alkaloids, heavy metals, cyanogenetic glucosides and santonin gave neg results, and the seed contained in each pill had all the morphological and anatomical characteristics of lucern.

A PAPINEAU COUTURE

Pyrethrum flowers from Kenya and Cyprus. ANON *Bull Imp Inst* 28, 425-9(1930) —Samples of pyrethrum grown in Kenya and in Cyprus were found to be of nearly as good quality as the best Dalmatian pyrethrum.

A PAPINEAU COUTURE

Tetrapone H BAGGESGAARD-RASMUSSEN *Archiv Pharm Chem* 38, 29-37(1931) —Mixts of alkaloid hydrochlorides have come into use in Danish hospitals. One contains morphine 50, narcotine 19.4, codeine 3.2, papaverine 4.5, thebaine 1.7 and narceine 0.8%. In another one thebaine and narceine are omitted and the narcotine is increased to 22%. The names "hexapone" and "tetrapone" are recommended for these as in line with Roche's "pantopone".

A R ROSE

Tinctura Quillajae and Liquor carbonis detergens. K A KARSMARK AND L KORLEK *Svensk Farm Tids* 34, 706-14(1930)(in German) —The tincture made with 62% alc contained 5 times more saponin than that made with 86% Quillajae tincture with 42% alc contained less sediment than the 62% prepn. The 62% alc prepn is to be preferred in making Liq carbonis detergens. Percolation is better than trituration.

A R ROSE

Sublimated or synthetic benzoic acid in preparing mucilago gummi arabici and sebum salicylatum. HUGO WASTENSON *Svensk Farm Tids* 34, 714-5(1930) —The next edition of the pharmacopoeia should specify synthetic $PhCOOH$ where it now specifies sublimated.

A R ROSE

The constitution and constants of sandal oil. R CAJOLA *Rivista stal essenze profum* 12, XXI(1930) —This essence is obtained by distg *Santalum album* wood. It is a yellow limpid liquid with an agreeable persistent odor and an irritating acid resin taste. It has sp gr 0.975-0.985 and boils and distils at 275-300°. The oil is sol in all proportions in 90% alc, and has 90-98% of α - and β -santalol, with α - and β -santalene ($C_{15}H_{18}$), santene ($C_{15}H_{16}$), santenone ($C_{15}H_{14}O$), sautenone alc ($C_{15}H_{14}O$) and traces of santalic acid ($C_{11}H_{12}O_7$) and terebantic acid ($C_{10}H_{16}O_4$) and a santal aldehyde ($C_{11}H_{12}O$). Sandal oil is used in medicine in mixt with other sedative and antiseptic substances, and

constitutes the principal astringent element in many pharmaceutical compounds. It is used in doses of 1.5-2 g. per day for treating the cure of hemorrhoids when other astringents cannot be used in the rectum because of the irritation of the mucosa. It is also useful in bronchial complications, as it is eliminated with the urine and during breathing.

R. SANSONE

The composition and properties of sassafras and its essential oil. R. CAJOLA *Kristall. essenze profumi* 12, XXII (1930). The wood of the sassafras plant contains the coloring principle sassafril and tannic acid. Its medicinal uses are listed. The roots contain balsamic resin, wax, sebaceous matter, tannic acid, coloring matter, gum, albumin, starch and 2-5% of a yellow essential oil. This oil is very little sol. in alc. it contains safrene (C₁₁H₁₆), an alc., safrole and a volatile phenol like substance.

R. SANSONE

The properties and composition of mustard seed. R. CAJOLA *Kristall. essenze profumi* 12, XXII (1930). Mustard seeds contain 20% of mucilage and 30% of a sweet fatty oil, the alkaloid sinigrin, the enzyme myrosin and sinigrin. Pounded mustard seeds, when beaten with lukewarm water develop a clam action between the myrosin and sinigrin producing glucose, KHSO₄ and allyl isothiocyanate. With boiling water the reaction does not take place. The uses of mustard seed oil are listed.

R. SANSONE

A study of the treatment for insomnia. HENRI FAVIER *Thesis*, Montpellier, 1929, *J. pharm. Alsace Lorraine* 57, 200 (1930). A monograph on the hypnotic action of sedormid (cf. C. A. 22, 3723).

S. WALDBOTT

Pharmacy in Syria. C. LADAKIS *Bull. Sud-Ouest et du Centre, J. pharm. Alsace Lorraine* 57, 213-4 (1930).—An account of present conditions under French administration.

S. WALDBOTT

"Virtual adrenaline" in relation to inactivation of adrenaline by formal. MARCEL PAGET AND CHARLES P. LEBLOND *J. pharm. chim.* [8], 12, 531-6 (1930), cf. C. A. 24, 3323-4.—When the NH function of adrenaline (A) is blocked by means of formal, the Denigès reaction becomes neg., but the reactions which are sp. for the pyrocatechol nucleus (Vulpian, Polin, P., cf. C. A. 25, 332) remain pos. Blocking the NH function diminishes considerably the mydriatic power of A without destroying it entirely. It sheds no light on the nature of "virtual A", work on this subject is in prepn. S. W.

Oxidation of official castor oil by potassium permanganate. Study of triazelaïn. G. SCHUSTER *J. pharm. chim.* [9], 13, 5-12 (1931).—The principle of the Hilditch method is applied (cf. C. A. 22, 244, 2177, 3702, etc.). By KMnO₄ oxidation in acetone soln. of official castor oil, 63.58% triazelaïn (A) was formed. After sapon of crude A, neither stearic nor dihydroxy stearic acid was obtained. The results of Heiduschka and Kirsten (C. A. 24, 1033) indicated these compds. were present. The same oil to which 3% of stearic acid was added yielded this quantity after oxidation; hence the sample examd. contained no glycerides of said acids. A was also synthesized by the action of epichlorohydrin of glycerol upon acid Na azelate (CO₂H(C₁₁H₂₁), CO₂Na) in sealed tubes at 135-140°. With a neutral 2% soln. of the Na salt of A, AgNO₃ forms a brick red ppt. not altered by light; solns. of BaCl₂, Bi(NO₃)₃, CdSO₄ and Sr(NO₃)₂ give a chamois-yellow, CuSO₄ gives a green ppt., MgCl₂ produces a chamois-yellow ppt. while Mg azelate is H₂O sol. By means of this reaction, A was purified.

S. WALDBOTT

Herba equiseti, a saponin-bearing drug. P. CASPARIS AND K. HAAS *Pharm. Acta Helv.* 5, 62-3 (1930).—The presence of saponin in the herb was proved by foam on expts. on a 10% chlorophyll free decoction and a 10% aq. soln. of a purified MeOH dry ext., and also by the hemolytic effect of 1% of the latter soln.

S. WALDBOTT

Senegin and its decomposition products. O. DAFERT AND E. KALMAN *Pharm. Acta Helv.* 5, 71-7 (1930).—By extg. senega root repeatedly with 96% alc. and evapg. to a small bulk, crude neutral saponin (senegin) is pptd. (yield 10%). By repeated pptn. from soln. in 80, 90 and 96% alc., 3-4% pure H₂O sol. senegin contg. C 52.29%, H 7.17% is obtained. It is hemolytic in diln. of 1:66,000. Hydrolysis with 5% H₂SO₄ and then 5% alc. H₂SO₄ for 30 hrs. yielded a substance identical with that examd. by Wedekind and Krecke (C. A. 18, 3362). Other products of hydrolysis are dextrose (41.09%), which seps. at first, then methylpentose (11.12%) and arabinose (11.06%) are split off (cf. Winterstein and Maxim, C. A. 13, 1213). A list of 21 references is added.

S. WALDBOTT

The evaluation of tribromophenol bismuth gauze. RUD. G. MAEDER *Pharm. Acta Helv.* 5, 112-6, 151 (1930).—Weigh 4-5 g. of gauze exactly, add 20 cc. of 25% HCl, knead the gauze with a glass rod, add 50 cc. hot H₂O, knead again, transfer the soln. into a 200 cc. flask and repeat the extn. with 5 X 30 cc. hot H₂O. Cool, filter off the sepd. C₆H₃(OH)Br₃ and in the filtrate ppt. Bi with H₂S. Collect the Bi₂S₃ on a small filter,

wash with 2×10 cc H_2O and dissolve the ppt. on the filter by adding, drop by drop, 15 cc HNO_3 (d 1.4). Wash the filter with 3×10 cc H_2O and warm on a steam bath to expel H_2S completely. Cool, then add in small portions 25 cc NH_4OH (20-25%), collect the $Bi(OH)_3$ on an ash free filter wash well, dry at $103-105^\circ$ and ash in a porcelain crucible. When cool, add a few drops of fuming HNO_3 , evap. on a water bath, then heat slowly with a flame, finally to red heat, cool in a desiccator and weigh. Multiply the wt. of Bi_2O_3 by 2 to obtain the wt. of "xeroform" in the gauze (com. xeroform yielded 49.6 and 55.2% Bi_2O_3 , 50% is the standard adopted). In 5 com. samples of gauze, marked 20%, 10, 10, 5, 5%, the method showed 16.0%, 6.4, 6.0, 4.37 and 2.95%, resp. Deviations from statements on the label must not exceed $\pm 10\%$. S. WALDBOTT

Pilulae ferratae Blandii. J. B. LANG *Pharm. Acta Helv.* 5, 121-3 (1930); cf. *C. A.* 24, 5110.—The pharmacol. results of Starkenstein (*C. A.* 21, 1494) confirm L's contention that Bland's pills should be as free as possible from ferric salt. The following modification of the previous test is given: Triturate 1 pill with 0.3 g $BaCl_2$, mix with 20 cc H_2O and filter. Divide the clear filtrate into 2 equal parts, to 1 add 5 drops of $K_4Fe(CN)_6$ soln., a deep blue ppt. is formed at once. To the other add 5 drops of $K_4Fe(CN)_6$ soln., at most a light blue coloration may appear. This test does not apply to pills consisting mostly of $FeCO_3$. S. WALDBOTT

Note on the Pharm. Ned. V. test for codeine in papaverine hydrochloride. E. ANSELER *Pharm. Acta Helv.* 5, 124 (1930).—The following requirement of Pharm. Ned. V. is probably based on error: "The soln. of 10 mg papaverine-HCl in 1 cc H_2SO_4 shall be colorless or at most faintly yellowish, upon adding 1 drop of $FeCl_3$ soln. and warming on the water bath, no violet color must form (codeine)." The violet color which at first is pale green is produced with papaverine itself, this fact was proved with samples of known purity, and with papaverine prep. synthetically. S. WALDBOTT

Determination of polysulfide and preparation of calcium sulfuraturn solutum (Vlemingx solution). J. BCCIN *Pharm. Acta Helv.* 5, 124-32 (1930).—The exact and rapid method of Schuef for the detn. of polysulfides (*C. A.* 19, 1829) is adopted in the new Pharm. Helv. V. This method enables a much needed control of the compn. of Vlemingx soln. (A). When prep. according to the directions of different pharmacopaeias including Pharm. Helv. IV, A will contain very variable quantities of polysulfides. To prep. A of const. compn., the formula adopted by Pharm. Helv. V is recommended: Slake 1 part CaO with 5 parts H_2O , then add 2 parts of washed S and 10 parts H_2O . Boil the mixt. for 1 hr. under reflux (to exclude the air), cool, decant and pour through cotton. The sp. gr. of A thus prep. is 1.12-1.14, the content of polysulfide S is 60-65 g. per l. A list of 8 references is added. S. WALDBOTT

The galenical preparations of Pharmacopoeia Helvetica V. K. STEGFRIED *Pharm. Acta Helv.* 5, 255-70 (1930).—An address. Among novel requirements demanded by the Pharm. of Swiss pharmacies are a standardized thermometer, an app. for evap. exts. under reduced pressure, special vessels to keep drugs and exts. over lime, a polarimeter and a polarization attachment to the microscope. Certain emulsions, e. g., of P, must be prep. with stabilized, s. e., enzyme free gum arabic. The new features of all galenical preps., e. g., the various exts., infusa, injectabilia, tinctures, ointments, etc., are discussed. S. WALDBOTT

The examination of alkaloids for purity in the Swiss pharmacopoeia. R. EDER, E. BÜRG AND H. T. LEX *Pharm. Acta Helv.* 5, 276-85 (1930), cf. *C. A.* 24, 5424.—An address. The various pharmacopaeias should agree on a uniform definition of alkaloids, e. g., in quinine sulfate. The reaction of alkaloidal salt solns. should be measured by detg. the pH value by means of colorimetric indicators. A list of 25 such salt solns. has been arranged into 6 groups of varying pH values (2.8-7.0). S. WALDBOTT

The use of *p*-hydroxybenzoic esters in sterilization and disinfection. TH. SABA-LITSCHKE *Pharm. Acta Helv.* 5, 286-8 (1930), cf. *C. A.* 22, 2029, 23, 2733, 24, 180, 3285, 4834, 5886, 25, 1031.—The antiseptic and sterilizing effect of 0.05% of nipasol (A), the Pr ester of $p-OHC_6H_4CO_2H$, exceeds that of a 0.25% $PhOH$ soln. E. Leschke found 0.04% of A or 0.2% of the sol. Na compd. completely non-toxic and non-irritating in injections. The 0.2% soln. is recommended for sterilizing instruments, and to keep medicaments sterile for internal therapy. A and the similar esters may also be used for disinfecting powders, e. g., $B(OH)_3$, talcum, warming with 1-3% Me ester, or 2-6% mixed esters to 70° , renders these powders germ free for at least 2.5 yrs. Addn. of these esters enhances the antiseptic effect of $B(OH)_3$ materially. S. WALDBOTT

A comparison of certain suspending agents. F. R. C. DATESON *Pharm. J.* 126, 52-3 (1931).—Shake a suspension of 2 g. prep. chalk in 100 cc. water for 30 min. with (a) 12.5 cc. of mucilage of gum acacia (*Brit. Pharm.*), (b) 12.5 cc. of mucilage of traga-canth, (c) 2.22 g. of compd. powder of tragacanth, (d) 3.33 g. of the same, then pour off

50 cc. of the suspensions and det. CaCO_3 in each suspension and residue. It is found that the most efficient suspending agent for insol. powders is *compt. powder of tragacanth* (Brit. Pharm.). The (above) proportion of 10 grains is but slightly less effective than 15 grains per fluid ounce, but as in the latter case the mixt. is very viscous, the weaker proportion is preferable. The 2 mucilages (a and b) are not nearly so effective as suspending agents. S. WALDRON.

The leaf oils of Washington conifers. Introduction. ARNOLD J. LEHMAN AND F. V. LANS. *J. Am. Pharm. Assoc.* 19, 840-3 (1930). — Botanists list 21 species of conifers in the state. The oils from 8 have been studied previously by others. The introduction contains a review of the literature and bibliography. L. F. WARREN.

Physiological potency of commercial ergot preparations. W. T. McCLOSKEY, M. R. THOMPSON AND N. G. HARBELLA. *J. Am. Pharm. Assoc.* 19, 844-7 (1930). A survey was made of the market supply of fluidext. of ergot and other ergot preps. The preps. were tested by the cockcombi method but if the results were not concordant checks were run by the Brown Clark method as modified by Pattee and Nelson. The fluidexts. assayed 0-400% of the U. S. P. requirements. The other preps., many of which were in ampoule form, assayed from 0 to about 120% of the labeled units. The assay methods used are capable of detg. the specific alkaloidal activity within $\pm 20\%$. Therefore those preps. assaying $\geq 120\%$ of the U. S. P. X potency may be considered as satisfactory. Of 79 com. fluidexts. 35% were substandard, 20% were of U. S. P. potency and 45% were superstandard. Many of the substandard preps. had been repackaged before final marketing. Probably some of the superstandard preps. had been assayed by the U. S. P. method immediately after making and the non-specific amine interference had instituted an error in standardization. Such errors in standardization may be avoided by aging the preps. in filled containers for several months before final standardization. Com. ergot preps. should bear an expiration date. L. F. WARREN.

Oleonephtha as absorber of ethereal oils from distillation water. V. I. VARENTZEV. *J. Chem. Ind. (Russia)* 6, 66-7 (1929). — Instead of following the usual procedure of returning the condensed dist. water holding droplets of ethereal oils back to the dist. app. the condensate collecting in the traps is made to pass through a layer of oleonephtha, which retains the ethereal oils. A saving of 1.77% of the product was accomplished by this innovation. This method was also successfully applied to the recovery of *terpineol* from condensate, increasing the recovered product by 3.33%. A. C. Z.

Ambergris and how to recognize it (MORRISON) 27. Analysis and preparation of medicated soaps (MAROLESA) 27. The economics of pyrethrum (GLASSFORD) 15. Pyrazolone derivatives [intermediates for drugs] (Ger. pat. 514,421) 25. Immunizing plants [vegetable serums for human and veterinary therapeutics] (Fr. pat. 693,723) 15. Creams, etc. (U. S. pat. 692,454) 27.

GILDEMEISTER, F., AND HOFFMANN, FR. *Die ätherischen Öle*. Band III. 3rd ed., revised by F. GILDEMEISTER. Mitteil. bei Leipzig. Verlag der Schimmel & Co. Akt. M. 36, half leather, M. 40. Cf. C. A. 24, 690.

TRUTWIN, HANS. *Grundriss der kosmetischen Chemie*. Brunswick. Vieweg. 152 pp. M. 9.75.

Synthetic drugs. I. G. FARBERVIND. A. G. (Erich Bartholomäus, inventor). Ger. 514,418, Apr. 28, 1928. Anthelmintics of a basic character are prepd. by treating ketotetrahydronaphthalenes with CH_3O and a secondary amine. Thus, an alc. soln. of 1 ketotetrahydronaphthalene, paraformaldehyde and $\text{Me}_2\text{NH}\cdot\text{HCl}$, boiled under reflux for 1 hr., yields 2-dimethylaminoethyl 1-ketotetrahydronaphthalene-HCl, m. 158-9°. Other examples are given also.

Synthetic drugs. I. G. FARBERVIND. A. G. (Hans Schmidt, inventor). Ger. 515,205, Nov. 7, 1928. Sol. complex org. compds. of Sb are prepd. by the reaction of aromatic α -dihydroxy compds. contg. acid salt-forming substituents or their salts, with oxygenated compds. of quinquivalent Sb. Thus, an aq. soln. of K catecholsulfonate may be treated with an excess of antimony acid and the soln. filtered and neutralized, the product is pptd. by pouring into MeOH. Other examples are given also.

Synthetic drugs. I. G. FARBERVIND. A. G. (Karl Streitwolf, Alfred Fehrlé and Hubert Oesterlin, inventors). Ger. 515,207, Aug. 28, 1929. Addn. to 510,451 (C. A. 25, 1036). The method of prep. arsenobenzenes having aromatic hydroxyacylamino groups described in Brit. 318,491 (C. A. 24, 2241), is modified by condensing one component in the form of its arsine with the other component in the form of its arsine oxide

or the corresponding dichloroarsine. Thus, 4 glycolylaminobenzenearsine oxide in alc. HCl soln may react with 3-amino-4-hydroxyphenylarsine. Other examples are given also.

Synthetic drugs. I. G. FARBENFABRIK A.-G. (Fritz Mietzsch and Heinrich Klös, inventors) Ger 515,466, Jan 30, 1929. New products effective against blood parasites are prep'd by acylating 4 (ω -aminoalkyl)aminobenzenearsonic acids which may or may not be substituted in the 2 position. The initial materials may be prep'd as described in Ger 501,088 (C A 24, 4588-9). Examples are given.

Synthetic drugs. SCHERING-KAHLBAUM A.-G. (Adolf Feldt, Walter Schoeller and Erich Borgwardt, inventors) Ger 513,799, May 6, 1929. Addn to 511,303. The method of Ger 511,303 (C A 25, 1335) is extended to the conversion of aminometal mercaptosulfonic acids generally into the corresponding ω -sulfomethylamino compds. The initial materials in two of the examples given are Na β -amino α -auromercaptoethanesulfonate and Na 3-amino-4-auromercaptoethanesulfonate.

Synthetic drugs. SCHERING-KAHLBAUM A.-G. (Walter Schoeller and Erich Borgwardt, inventors) Ger 514,506, Feb 12, 1926. Addn to 506,443. The method of prep'g auromercapto acids described in Ger 506,443 (C A 25, 381) is modified by using acid or neutral sulfite in place of H_2SO_4 . Examples are given.

Synthetic drugs. KARL LÖBECK & CO. Ger 515,092, Sept 11, 1928. Salts of choline with negatively substituted aromatic arsonic acids are prep'd by the customary salt-forming methods. The products are cryst and non hygroscopic. Examples are given describing the prep'n of the choline salts of, *inter alia*, 4-iodo-3-nitrobenzene arsonic acid, 3,5-dinitro-4-aminobenzene arsonic acid, and 4-carboxybenzene arsonic acid.

Extraction and concentration of vitamins. G. DUBOIS Belg 372,102, Aug 31, 1930. Ground fish ovaries and livers are treated in strong alc. The oily liquid obtained on pressing is saponified with an alc. alkali soln, and the soap is treated with $MgCO_3$ and H_2O to obtain vitamin A and with superheated steam to obtain vitamin F.

Medicinal silver preparation. PRIVILEGIUM-CHIMISCHES LABORATORIUM HUGO ROSENBERG Ger 515,061, Oct 18, 1927. An ammoniacal soln of Ag_2O is heated with an excess of starch until the oxide is wholly or partly reduced. An example is given.

Antiseptic solution. THE PERSOENT CO. Fr 693,083, Mar 31, 1930. An antiseptic and disinfectant soln contains, e. g., chlorothymol 0.1, glycerol 7.5, EtOH 25, AcOH (of 36%) 1 and water 64.4%.

Creams, pomades, etc. COMPAGNIE FRANÇAISE POUR L'EXPLOITATION DES PROPRIÉTÉS THOMSON-HOUSTON Fr 692,757, Mar 26, 1930. A protective pomade or cream contains soap 19, gum arabic 4, lanolin 2, glycerol 1 and water 74 parts. Gum arabic may be replaced by dextrin and lanolin by other oily substances.

Perfumes. J. D. RIENEL & DE HAEN A.-G. Ger 515,332, Dec 8, 1926. Ethers of 3,4-dihydroxybenzyl alc. are used as fixing agents or solvents for perfumes. The prep'n of the dimethyl, methyl ethyl and diethyl ethers is described.

Essence d'Orient. JOSEPH JACOBSON Fr 693,735, July 22, 1929. Pure Essence d'Orient is obtained by emulsifying the org. materials mixed with it by peralcohol which permits a rapid decantation from the crystals.

18—ACIDS, ALKALIES, SALTS AND SUNDRIES

E. M. SYMMES

Causes of the intoxications produced in the cleaning of sulfuric acid tanks and tank cars. LEYMAN and WEBER *Zentr. Gewerbehyg. Unfallverhütung* 17, 154-9 (1930), *Chimie & Industrie* 24, 832-3 (1930).—From a study of a large no. of cases it is concluded that the intoxication is due to AsH_3 . Against this were the suppositions that the action of H_2SO_4 on metallic Fe in the presence of reducible As compds. does not produce AsH_3 , and that, as com. H_2SO_4 (except contact process H_2SO_4) contains HNO_3 , AsH_3 gas cannot form. Lab. expts. showed that when crude dil. H_2SO_4 acts on Fe in the presence of reducible As compds., AsH_3 can be evolved, even at atm. temp., at higher temps., even when very dil. H_2SO_4 can still give rise to evolution of AsH_3 under similar conditions, finally, the small quantities of HNO_3 present in technical H_2SO_4 do not prevent the reaction. When cleaning out the tanks or cars, the H_2SO_4 which sits the sludge heats up on being dild., attacks the Fe and ultimately causes evolution of AsH_3 . To prevent the intoxications, cleaning outlets should be provided in the bottom of the tanks, or the washing should be carried out with a large amt. of water, which is continuously siphoned off.

A. PAPINEAU COUTURE

Lime in 1929. A. T. COONS. *Bur. Mines, Mineral Resources of the U. S.* 1929, Pt. II, 267 (88 (preprint No. 20 published January 14, 1931)). R. H.

Density of sulfuric acid solutions of copper sulfate. G. CHAPUIS, C. CHARMETANT AND A. RAU. *Chim. et Ind.* 24, 791 (1930). Control of the amount of $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ requires frequent dilution of the compound of the acid to maintain the acidity within relatively narrow limits. As the dilution of acidity is relatively long to carry out, curves of the compound of acid at different densities are calculated and plotted as triangular coordinates, the sides of the triangle representing CuSO_4 and H_2O respectively and from them the crystal curve was obtained. By anamorphosis the crystal curve was transposed to rectangular coordinates. The diagrams permit obtaining the compound of H_2SO_4 , CuSO_4 , and H_2O from the density and CuO contents. For the curves and method of using them the original article should be consulted.

The sulfates in industry: a review of the water-soluble products. IV. Magnesium sulfate. V. Iron sulfates. VI. Zinc sulfate. VII. Copper sulfate. CHAS. H. BURCHER. *Chem. Ind.* 1, 88, 124 (1931) at 1, 25, 127. An outline of characteristics required of these sulfates in the industries. W. H. ROBERTSON.

The problem of the drying of salt. H. L. LARSEN. *Chim. et Ind.* 24, 805 (1930). A brief discussion is given of various types of apparatus for drying salt from a moisture content of 3.8% down to 0.5%, together with approximate figures on the cost of the drying.

Studies on fundamental synthesis of calcium aluminates and their hydration. VII. SHUICHIRO NAGAI AND RYUICHI NAITO. *J. Soc. Chem. Ind. Japan* 31, 5 (April 1930), 315 (1930) at 1, 24, 470. Tested on that $\text{CaO} \cdot \text{Al}_2\text{O}_3$ and $\text{CaO} \cdot \text{Al}_2\text{O}_3 \cdot \text{H}_2\text{O}$ can be obtained by heating appropriate mixtures of CaO , Al_2O_3 , and SiO_2 above 1200°C for a longer time at 1400°C . In the mixture between CaO aluminate and CaO silicate, the combination of CaO and Al_2O_3 or SiO_2 is completed instantly by heating at a little higher temperature than 1900°C . Compression tests on the synthesized products of alite and C_3S type and of eutectic mixtures show that the alite has a small value at about curing times, but the strength increases gradually on longer curing in water. The C_3S type has no perceptible strength, but its strength increases remarkably when 5–10% of CaO is added.

Anhydrous aluminum chloride: its uses and manufacture. C. SWAN. *Chim. et Ind.* 24, 1817, 21 (1930). A brief review is given of the more important uses of AlCl_3 with an outline of the chief processes of manufacture, more particularly a comparison of the MacFarlane and Haggard processes.

Chile's attempt to modernize and make profitable the production of nitrate. L. NIELSEN. *Austral. J. Sci.* 27, 810 (1930) – Statistics. C. J. S.

Annual review number. *Eng. Mining J.* 191, 97 (1931). GOLD. ROBERT J. GRANT. *Ind.* 97 (1931) – Increase in mining types the period of economic adversity. SILVER. H. M. BRATTON. *Ind.* 98 (1931) – The by-product output of the metal makes its future uncertain. Precious stones. GEORGE F. KENT. *Ind.* 99. LEAD. CLAYTON H. CRANE. *Ind.* 100–1 – Abnormal refined metal stocks are on hand. Asbestos. NORMAN R. FISHER. *Ind.* 101–2 – Russian output has rapidly increased, while that of the U. S. A. and the Union of South Africa has fallen markedly. Cobalt. G. C. RATTEN. *Ind.* 102 – The ceramic industry consumes about 40% of the output, but increased use in Co -alloys is noticed. Aluminum. C. I. MANTRELL. *Ind.* 103 – The Al industry has held its own, with an outlook for further expansion in 1931. Copper. ARTHUR NORMAN. *Ind.* 103–4 – Opposition to law of supply and demand has again proved futile. Graphite. BENJAMIN L. MURRAY. *Ind.* 104. Zinc. JULIAN D. COOPER. *Ind.* 105 & Nickel. THOMAS W. GILSON. *Ind.* 106. Iron. EDWIN C. TUCKER. *Ind.* 107. Tin. H. RALPH SCOTT. *Ind.* 108 (1931) Sulfur and pyrites. RAYMOND P. RACON. *Ind.* 109. Quicksilver. H. W. GORTON. *Ind.* 110. Minor metals. PAUL M. TYLER. *Ind.* 111–4 – Economic figures for Sb , chromite, lanthanum, Mn, Mo, W, U, Ra, V and Ta are discussed. Non-metallurgy. OLIVER BOWLER. *Ind.* 115–8 – Sales of construction materials led the decline in the depression period. Barites, feldspar, fluor spar, gypsum, kaolin, magnesite, phosphate, mica, talc and soapstone and silica are discussed. Platinum. H. W. ROBERTSON. *Ind.* 118 – The decline in price during 1930 was due more to increased production than to marked decrease in consumption. Increased consumption was evident in the essential industries. Uncommon ores and metals. H. C. MYER. *Ind.* 119–20. Elements discussed are: Be, Ca, Rn, Ce, Se, Fe, Th, Ti and Zr. World potash industry prospers. J. W. TUCKER. *Ind.* 121. Mining in the United States. S. D. STRAUSS. *Ind.* 121–4. Canadian output of metal expands. S. J. COOK. *Ind.* 123 & Mining in South Africa and Rhodesia. OWEN LEITCH. *Ind.* 126–7. Mining in Latin America in 1930. S. D. STRAUSS.

Ibid 128-30 Europe, Siberia, Far East and India. DONALD J. PARQUHARSON *Ibid* 131-2 The mining industry of the U. S. S. R. V. KAMASHOV. *Ibid* 132-3. Mining in Australasia in 1930. PETER G. TAIT. *Ibid* 133. Japan. SHIRONAO NISHIMURA. *Ibid* 134. W. M. BOYNTON

A study on agar-agar. H. TETSUNOBUKE YAMAGAWA and Y. NISHIDA. *Repts Imp Ind Research Inst, Osaka, Japan* 11, No 14(1930), cf C A 23, 4538—Constants of *Gelidium amansii* Lamx and the phys. properties of agar-agar are given. F. I. NAKAMURA

Study of the conditions of cooking of gum tragacanth. A. T. LOMANOVITCH and V. V. PETUKHA. *Izv Tekstiln Promysl* 9, No 3, 71-3(1930) *Chimie & Industrie* 24, 147(1930)—From a study of the various factors involved, the following procedure is recommended: swell the finely divided gum in 20 parts of water at 50° for 12-15 hrs with continuous stirring, heat for 2 hrs. in an open kettle and cool with stirring for 6-8 hrs., at which point max. viscosity is attained. A. PAPINEAU-COUTURE

The production of superphosphate from Khibinsk apatite (VOL'KOVICH, *et al*) 15.

British Plastics Year Book, 1931. London: The British Plastics Moulding Trade Assocn 229 pp

Hydrocyanic acid. IMPERIAL CHEMICAL INDUSTRIES, LTD, T. S. WHEELER, J. McALLAN, W. B. FLETCHER and H. A. T. MILLS. *Brit* 335,595, March 27, 1929 *Fr* 692,542, March 21, 1930 Reaction between a hydrocarbon gas (such as $C_{12}H_{10}$, C_6H_6 , and various natural and artificial industrial gases contg hydrocarbons) and NH_3 to produce HCN and H_2 is effected by passing a mixt. of the hydrocarbon material with more than one mol. proportion of NH_3 for each at. proportion of C, at a temp. of at least 1150°, rapidly through an unpacked reaction chamber. H_2 or water vapor may be present, the gas mixt. may be preheated to 600° or higher, and the reaction chamber is preferably formed with smooth internal surfaces such as glazed silica or sillimanite. H_2 compds. and CO_2 may be removed from the gases by passing them through aq. or liquid NH_3 .

Hydrocyanic acid. IMPERIAL CHEMICAL INDUSTRIES, LTD. *Fr* 692,543, March 21, 1930 HCN is produced as described in the preceding patent, but the NH_3 is not used in excess. Instead an appreciable amt. of a diluent such as H_2 or N_2 is used. The hydrocarbon may be coke furnace gas, and it may be purified and at the same time mixed with NH_3 by bringing it into contact with liquid NH_3 .

Phosphoric anhydride. L'EVENCE COPPEL. *Belg* 372,202, Aug 31, 1930 Phosphates, to which may have been added a substance capable of reducing the reaction temp., e. g., Al_2O_3 , are heated in presence of SiO_2 .

Phosphoric acid. H. LAWASSÉE. *Belg* 370,074, June 30, 1930 Pb oxychloride or hydroxide is used in a closed cycle to isolate phosphoric acid as Pb phosphate in a salt soln. contg phosphoric acid.

Phosphoric acid; hydrogen. BAVERISCHE STICKSTOFFWERKE A.-G. (Vladimir N. Ipat'ev, inventor). *Ger* 514,173, Mar 1, 1929. See *Fr* 670,338 (C A 24, 1941)

Phosphoric acid and aluminum compounds. I. G. FARRINGTON A. G. *Fr* 693,818, Apr 12, 1930 Al is pptd. as $AlCl_3$ by HCl gas from a hydrochloric acid soln. of substances contg Al and H_2PO_4 , and the H_2PO_4 is recovered after elimination of the Al . An example is given of the treatment of natural Al phosphate.

Silicic acid. RICHARD WILLSTÄTTER, KARL LOEWINGER and HAINRICH KRAUT. *Ger* 514,414, Nov 3 1928 Stable solns. contg silicic acid as simple mols. are prepd. by decomp. an alkali silicate soln. with an acid, or by the reaction of a silicon halide with an aq. suspension of Ag_2O or other suitable base, and acidifying the soln. so obtained until the hydron concn. lies between pH 1.5 and 3.5. Examples are given.

Apparatus for effecting gas reactions such as catalytic production of sulfuric acid, ammonia or methanol. LUCIO CASALE. U. S. 1,790,853, Feb 3. Various details are described of an app. with baffle plates and partitions arranged for heat-exchange between products of reaction and gases supplied for the reaction.

Removing arsenic from acids. METALLGES A.-G. *Fr* 692,595, Mar 24, 1930 As is removed from a concd. acid such as H_2SO_4 of 60° Bé by the addn. of the exact amt. of H_2S necessary to convert the As to sulfide.

Purifying acids or acid salts. PAUL BOESSNECK. *Ger* 514,499, Sept 26, 1929 Solns. of acids or acid salts obtained by treating solns. of Ca salts with H_2SO_4 are freed from dissolved $CaSO_4$ by addn. of HIP .

Salts of cyanic acid. DEUTSCHE GOLD- UND SILBER-SCHNEIDANSTALT VOM ROESSLER. Fr 692,481, Mar 21, 1930. Salts of cyanic acid are prep'd by heating, in the absence of solvents, urea with alk. earth compds. such as CaO , Ca(OH)_2 , CaC_2 or alkali compds. such as Na_2CO_3 or K_2CO_3 .

Synthesis of ammonia in presence of inert gases. AMMONIAQUE SYNTHETIQUE ET DERIVÉE, SOC. ANON. Belg 372,037, Aug 31, 1930. After having been freed from NH_3 , the gases are partly returned to the same tower along with the gas from the preceding tower (or the fresh gas in the case of the 1st tower), and partly sent to the next tower (or the gas discharge in the case of the last tower).

Synthesis of ammonia. SOC. CHIM. DE LA GRANDE PAROISSE (AZOTE ET PRODUITS CHIM.) Fr 693,712, July 9, 1930. Small quantities of CO_2 in the gases for the synthesis of NH_3 are removed by the action of the CO_2 , NH_3 , and steam to form $(\text{NH}_4)_2\text{CO}_3$, sufficient steam being left in the gas, so that during the final condensation it carries down the $(\text{NH}_4)_2\text{CO}_3$.

Gaseous syntheses. UNIVERSITÄTS-INSTITUT FÜR PHYSIKALISCHE GRUNDLAGEN DER MEDIZIN. Fr 693,641 Feb 28, 1930. In gaseous syntheses of surface reaction type by means of elec. discharges, the reaction chamber in which the elec. discharges are passed through the gas is heated to above the critical temp. of the gas to be prep'd. Examples mentioned are the preps. of NH_3 and C_2H_2 .

Oxidation of ammonia. E. I. DU PONT DE NEMOURS & Co. Fr 692,824, Mar. 27, 1930. NH_3 is oxidized to oxides of N by passing a mixt. contg. NH_3 and O in contact with a catalyst of the Pt group, the exposed surface of which per unit section of the gaseous current is more than 4 thicknesses of gauze of 23-63 mesh per linear in., formed of threads of 127-25 thousandths of a mm. in diam.

Oxidizing ammonia. E. I. DU PONT DE NEMOURS & Co. Fr 693,559, April 8, 1930. NH_3 is oxidized to oxides of N by passing a mixt. of NH_3 and O over catalysts and rapidly cooling the reaction mixt. to below 500° . An app. is described.

Oxidizing ammonia with oxygen. FRANK G. LILJENROTH. Ger 514,392, Sept. 9, 1926. See Swed. 63,200 (C. A. 24, 1473).

Alkali ferrocyanides. HAMBURGER GASWERKE G. M. B. H. Fr 692,868, Mar. 28, 1930. Alkali ferrocyanides are prep'd during the extrn. of HCN from coal gas by washing, by means of bivalent compds. of Fe suspended in an alk. lye contg. an excess of alkali. The excess of alkali depends in amt. and concn. on the content of HCN in the coal gas, so that only sol. ferrocyanides are produced, and this excess after being transformed into carbonate or bicarbonate by the CO_2 in the gas is neutralized.

Alkali chromates. METTAL CHEMICAL CO. OF AMERICA. Fr 692,614, March 24, 1930. See U. S. 1,732,863 (C. A. 24, 2533).

Alkali or alkaline earth salts from the corresponding chlorides. H. LAWARRÉE. Belg. 370,808, July 31, 1930. The acid is brought into contact, in aq. medium, with the chloride in presence of an oxide, hydride or basic mercurous or mercuric salt which can combine with the Cl ion. The HgCl_2 or Hg_2Cl_2 produced is treated with a base to regenerate the oxide or other mercurous or mercuric salt. Cf C. A. 25, 779.

Chromium compounds. GINO PANEBIANCO and VIRGINIO A. DE LIBERA. Fr 692,786, Mar. 26, 1930. Industrially pure Cr compds. particularly a mixt. of Cr_2O_3 and H_2CrO_4 or its salts, are obtained from crude materials, such as Cr ores or slags, by mixing the crude material with substances contg. pure or combined C and submitting the mixt. to the catalytic action of Cl in the presence of free or combined O.

Alkali metal phosphates. METALLGES. A-G. Brit. 335,492, Oct. 17, 1929. Phosphates such as Na_3PO_4 and K_3PO_4 are produced by fusing (suitably at a temp. of about $1150-1250^\circ$) a mixt. of alkali metal sulfate and a metal P alloy such as ferro-P, P-Cu or P-Mn (preferably in the form of powder or small lumps) in a elec. resistance furnace. Various details of procedure for avoiding atm. oxidation, etc., are described. The reaction product is leached with water, and metal sulfide is obtained as a by-product.

Phosphates. PHOSPHATE RECOVERY CORP. Fr 692,731, Mar. 23, 1930. Phosphatic material contg. particles of phosphate and particles of gang is conc'd. by treatment with alkali and with oil and sepn. by a surface flotation. An app. is described. Cf C. A. 24, 1473.

Phosphates. SOC. D'ÉTUDES POUR LA FABRICATION ET L'EMPLOI DES ENGRAIS CHIMIQUES (Etienne T. de Boismilon, inventor). Fr 693,735, July 22, 1929. A Ca phosphate having a content of citrate-sol. P_2O_5 of at least 35% of the dry product is obtained by adding the ground natural phosphate very slowly to the HCl soln. (contg. not more than 70 g. HCl per l.) at a temp. of about 30° .

Phosphotungstates. JOSSE & F. DARLING (to E. I. du Pont de Nemours & Co.)

U S 1,700,905, Feb 3 For purifying phosphotungstates such as Na phosphotungstate in a mother liquor the crude tungstate liquor is acidified (suitably with HCl), an aromatic amine such as dimethylamine is added, the reprecipitated phosphotungstate formed is removed, the ppt is treated with an alk. substance such as NaOH and the amine is steam distilled off.

Metal carbonyls. I. G. FARBERNINO A-G (Leo Schlecht and Emil Keunewecke, inventors) Ger 515,464, Aug. 19, 1928. See Fr. 679,542 (C. A. 24, 2865).

Manufacture of ammonium sulfate from the spent gases of coke ovens and synthetic-ammonia plants. R. LEDENT Belg 370,942, July 31, 1930. FeSO₄ is used in the neutralizers instead of H₂SO₄. The pptd. Fe(OH)₃ is air-dried and used as Fe ore. The concd. (NH₄)₂SO₄ soln is crystal.

Utilizing dilute sulfuric acid for the manufacture of ammonium sulfate. Soc. GÉNÉRALE DE FOURN. À COKE (SYSTÈMES LECOQ) Fr 692,789, Mar 26, 1930 See Belg 359,553 (C. A. 24, 850).

Aluminum chloride. I. G. FARBERNINO A-G (Johannes Brode and Carl Wurster, inventors) Ger 515,033, June 2, 1929 Anhyd. AlCl₃ free from Fe is prepd. by fusing the crude salt with a reducing agent, e. g., Al, and sepg. Fe from the melt by electromagnetic means, e. g., by suspending an electromagnet in the melt.

Magnesium chloride. ELISA LACELL NÉE VANDER-VORDE Fr 693,232, April 3, 1930 Carnolite or MgCl₂ is dehydrated by adding gradually to a relatively large anhyd. mass of MgCl₂ or MgCl₂ and KCl relatively small quantities of the salt to be dehydrated, to cause a rapid evapn. of the water and prevent the formation of oxides.

Potassium nitrate. KALI-FORSCHUNGS-ANSTALT G. M. B. H. Fr 693,181, April 2, 1930 A mixt. of Al(NO₃)₃ and KCl is heated until no more nitrous gases or Cl sep., and the KNO₃ is dissolved in water or in KNO₃ soln. The Al₂O₃ is reconverted to Al(NO₃)₃ by HNO₃, the nitrous gases being used as well. Cf. C. A. 24, 3091.

Potash and alumina. FELIX JOURDAN Fr 693,074, Mar 31, 1930 KOH and Al₂O₃ are obtained from leucite by heating limestone with leucite and a suitable flux to the m. p., cooling, grinding and leaching with a soln. of K₂CO₃ to neutralize the CaO and ppt. it as CaCO₃.

Selenium sulfide. COMPAGNIE FRANÇAISE POUR L'EXPLOITATION DES PROCÉDÉS THOMSON-HOLSTON Fr 692,758, Mar 26, 1930 Stable active SeS₂ of definite compn. is made by satg. a soln. of AlCl₃ with H₂SO₄ and adding a soln. of H₂SeO₃ with continued addn. of H₂SO₄. A yellowish sol. is first formed which is converted to a yellow flocculent ppt. A temporary and moderate excess of acid gives a sol. The ppt. is washed and dried. BaCl₂ may be used instead of AlCl₃. Another method consists in mixing S and Se in the theoretical proportions and heating to 225°. The plastic product obtained slowly changes and becomes fragile and hard. The SeS₂ may be combined with a hydrophile medium such as bentonite by heating the 2° to 125-150°. The product may be used for the vulcanization of rubber or in aq. suspension as a fungicide or insecticide.

Sodium cyanide. S. COCLIER Belg 372,147, Aug 31, 1930 A mixt. of earthenaceous material, alkali salts and catalyst is treated with hot N₂ with addn. of a quantity of Na salts equiv. to the quantity of NaCN required. After sepn. of NaCN, the mass is treated with steam under pressure to produce formates and NH₄. Cf. C. A. 25, 385.

Recovery of sodium carbonate from "tails salts." HARRY W. MORSE U S 1,791,281, Feb 3 "Tails salts" contg. Na₂CO₃ such as those produced from Searles Lake brine, are washed with water at a temp. of 30-40° in order to dissolve out Na₂CO₃.

Sodium chromate, alumina. I. G. FARBERNINO A-G (F. Wissing, inventor) Ger 513,942, Feb 5, 1930 Addn. to 505,318 In the method of Ger 505,318 (C. A. 24, 5949), the extn. of the sol. portion is effected by heating the melt with water under pressure to a temp. above 100°, e. g., to 230°. The ext. thus obtained is relatively free from SiO₂.

Sodium fluoride. E. V. BAITZKE, W. I. DRENFELL and M. E. JAKUBOWITZ Brit 735,703, Aug 29, 1929 NaF (neutral, acid or mixed) is obtained by the action of gaseous HF on solid calcined soda at an elevated temp. (suitably 70°), and the acid fluoride formed may be further neutralized with calcined soda. App. is described.

Acid disodium pyrophosphate. METALLGES. A-G Brit 335,591, May 17, 1929 This compd. is obtained by evapn. *in vacuo* (suitably at 180-200°) a soln. contg. H₃PO₄ (which may be that obtained by decompg. crude phosphates with H₂SO₄) and a proper proportion of NaCl.

Bleaching powder. I. G. FARBERNINO A-G Ger 514,390, May 24, 1928. Ground Ca(OH)₂ is suspended in an inert liquid, e. g., CCl₄, and treated with Cl at a temp. above

40°, the mixt. being rapidly cooled below 20° when the reaction is completed. The product is dried *in vacuo* at a low temp.

Mixtures of hydrogen and carbon oxides. BRITISH CELANESE, LTD., W. BADFA and E. F. STIMPSON. Brit. 335,632, July 2, 1929. Mixts. of CO and H₂ such as those obtained by releasing the pressure upon wash waters used in treating industrial CO and H₂ mixts., have their H₂ content reduced or eliminated by treatment with chromite above 700° (preferably at 800–900°) which increases the CO content of the mixt. so that it may be used for CO production by drying and passing it over hot coke or charcoal.

Animal black. BAUGH & SONS CO. Fr. 693,239, April 3, 1930. See Ger. 512,484 (C. A. 25, 1045).

Roasting sulfur ores. SAUREFABRIK SCHWAFERHALL. Fr. 692,660, Mar. 22, 1930. Sores or other materials contg. S are roasted in mch. furnaces of several floors in countercurrent to air, the ore, etc., passing over intermediate floors to which only part of the air has access to prevent premature cooling by the air.

Condensation products. I. G. FARBENIND. A. G. Fr. 693,750, April 11, 1930. Condensation products are obtained as a fine powder without odor by the reaction of aliphatic aldehydes, or their mixts. with bases of the naphthalene series in the presence of acids in aq. sol. in water and treatment of the products with water. Examples are given.

Condensation products of urea and formaldehyde. MARTIN LUTHER and CLAUS HETCK (to I. G. Farbenind. A. G.). U. S. 1,791,062, Feb. 3. Condensation is effected with an aq. soln. of urea and CH₂O in which a pH of 4 to 7 is maintained by addn. of buffer compds., and there is added to the condensation product, prior to hardening, a substance such as MgCl₂ or formamide which is capable of forming acid in the liquid condensation product during the heat treatment employed for the hardening.

Resitols. AGG. NOWACK A. G. and RICHARD HESSEN. Fr. 693,004, Mar. 29, 1930. The initial sol. and fusible condensation products (resols) from phenols and CH₂O, if necessary with the addn. of accelerators or the polymers or homologs of, CH₂O, are heated in thin layers for a short time, preferably repeated, until the whole or a greater part of the volatile constituents are eliminated and the resitol formed contains the desired amt. of resol for the final treatment of the resulting resitol.

Cyanamide-formaldehyde condensation products. HANS SCHMIDT (to I. G. Farbenind. A. G.). U. S. 1,791,433, Feb. 3. Formaldehyde is caused to act on cyanamide in aq. soln. contg. a "strong" inorg. acid such as dil. HCl, while heating (suitably at a temp. of the water bath) until the condensation product sepa. in an amorphous condition. U. S. 1,791,434 also relates to effecting reaction between cyanamide and CH₂O, as by heating in dil. aq. soln.

Plastic materials. COMPAGNIE FRANÇAISE D'ACCUMULATEURS ÉLECTRIQUES. Fr. 692,565, Mar. 22, 1930. A plastic material for molding hard elastic products, particularly suitable for molding accumulator vats, contains a mixt. of cotton-oil tar from the residues obtained in purifying cotton oils with H₂SO₄, petroleum tar, fibrous materials and a mineral charge. An example contains cotton-oil tar 40, petroleum tar 20, kieselguhr 30 and asbestos or jute 10%.

Molding casein. ADALBERT ZSICHMONDY. Ger. 515,376, Dec. 4, 1926. See Brit. 298,815 (C. A. 23, 3061).

Composition for use in modeling, sculpture, pottery, etc. KARL F. HÖLLER and STEFAN SCHLESS. Ger. 515,442, Dec. 28, 1928. See Austrian 119,622 (C. A. 25, 784).

Stabilization of dispersions of substances little soluble in water. N. V. BATAAFSCHE PETROLEUM MAATSCHAPPIJ and HYMAN LIMBURG. Dutch 23,111, Dec. 15, 1930. Aq. emulsions or suspensions of water-insol. substances, oils, tar, asphalt, rubber, etc., are stabilized by the addn. of small quantities of sulfonic acids. The latter are derived from petroleum products (lubricating oil fractions) by repeated sulfonation with 10% oleum (20% SO₃) at 50°. The acid sludge from the 4th or 5th sulfonation is neutralized, and the sulfonic acids are dissolved in water. In a 5% alkali sulfonate soln. 160% of asphalt can be dispersed to a stable emulsion. The active sulfonic acids have a min. mol. wt. of 300, are insol. in xylene and CCl₄, and are not pptd. by CaCl₂. Cf. C. A. 24, 444.

Adhesive for belts and the like. A. FLOWSKIER. Belg. 369,970, June 30, 1930. The product is a mixt. of lanolin, wax (e. g., paraffin) and fish oil (e. g., cod liver oil).

Adhesive. WILLIAM B. WESCOTT (to The Rubber Latex Research Corp.). Can. 308,567, Feb. 10, 1931. A plastic adhesive for laminated articles comprises heat-coagulated hemoglobin, rubber and a vulcanizing agent. It is used for securing rubber and rubber compns. to metal, bakelite, wood, etc. Cf. C. A. 24, 2257.

Wetting, etc., agents. H. TH. BÖMMER A. G. Fr. 692,862, Mar. 28, 1930. Wetting

19—GLASS, CLAY PRODUCTS, REFRACTORIES AND ENAMELED METALS

G. E. EASTON, C. H. KERR

Glass, its composition and properties. GEO W MOORE *J Chem Education* 8, 421-41(1931)—A review F. H.

The theory of the composition of glass. E. ZSCHUMMER *J Soc Glass Tech* 14, 229-49T(1930)—Graphic methods of representation of the compn. of glass as a quaternary system are discussed H. F. K.

Effect of alumina and silica on some properties of glass. M. A. BESBORODOV *Glashtite* 60, 5-9(1930)—An increase in the Al_2O_3 and SiO_2 contents retarded the devitrification of a $B-SiO_2-Al_2O_3-CaO$ potash-soda glass, and increased the resistance toward water and Na_2CO_3 . The resistance toward $NaOH$ and H_2SO_4 rises with an increase in the SiO_2 content but falls with an increase in that of Al_2O_3 B. C. A.

The influence of manganese oxide on some properties of glass. A. A. CHILDS, VIOLET DIMBLEBY, F. WINKS AND W. E. S. TURNER *Ceramic Industry* 15, 292(1930)—The replacement of Na_2O by MnO in 3 series of glass batches resulted in lowering the working temp and the coeff of thermal expansion; it raised the annealing temp and improved the chem stability M. A. F.

Feldspar, its effect in glass. C. W. KORTIG *Ceramic Age* 16, 150-2(1930), *Glass Ind* 11, 261-6—Feldspar in the glass batch is considered from the standpoint of the beneficial effects of alumina. Feldspar causes no mech or chem disturbance in the batch, is non-corrosive to tank blocks, and hinders surface sepn of silica. It should be added at intervals in small quantities to avoid producing glasses of different compns in the same tank. A bibliography of 19 titles is given M. A. E.

Plea for uniformity of glass-making materials. R. L. FRINK *Natl Glass Budget* 46, No. 26, 3, 13, 16, 20-1, 21(1930)—P. discusses at length the undesirable effects of variation of Fe content and moisture content, also of the presence of extraneous reducing agents in the raw materials of the batch, sometimes including not only tar and soot, but also catsup, sirups and fruit juices introduced with bottle cullet. Checks and seeds as well as unsatisfactory color, brittleness, and lack of homogeneity in glass may all be traced to such variations in materials M. A. E.

Effect of various radiations on the coloring of glasses. P. GILARD AND AD. LECRENIER *Chimie & Industrie* 24, 1035-51(1930), cf *C. A.* 21, 3434, 25, 1049—A large no. of glasses of known compn. were exposed for several weeks to the light of a Hg-vapor arc. The same and other glasses were exposed to the action of Ra rays for 29 days, and also samples of minerals and of a no. of pure chem compds. The absorption spectra of the irradiated glasses were measured by the method of G. and Swings (*C. A.* 24, 5235). From the results G. and L. are inclined to the belief, contrary to Przibram and coworkers, that the action of the Ra rays is like that of the ultra-violet rays in that it affects the mol rather than the atom A. PAPIEUAU-COUTURE

Note on the influence of different radiations on the coloring of glass. AD. LECRENIER AND P. GILARD *Rev universelle mines* 4, 349-62(1930) cf *C. A.* 25, 1049—The ultra-violet ray has a marked influence on the intimate compn of glass. It provokes reactions that may modify the aspect of the glass. The fact that certain glass seems unaltered does not mean, necessarily, a lack of internal reactions; the existence of these can be established only through sensible phenomena, such as colorations. In glass containing Cr or other elements existing in several states of oxidation, the action of the ultra-violet ray is clearly shown by displacement of color. The color obtained varies essentially with the compn. of the glass. Because of the lack of permeability of glass to ultra-violet rays, the color is at first superficial and only with time reaches the deep layers and the center of the mass. The color obtained disappears under the action of heat near the softening temp. of the glass. The heat causes reactions that are the reverse of those produced by the ultra-violet ray and destroys the mol structure the latter has established. The ultra-violet ray acts only chemically and modifies only the mol structures of bodies. The Ra ray causes deeper disintegrations and attacks the atoms; however, it is possible that under the action of Ra rays there is produced a modification in the relative positions of atoms in the mol. Tests with Ra rays have shown that all glass contg K_2O gives yellowish green shades while that contg Na gives shades of brown, the coloring varying with the compn. of the glass, often near the softening temp. Glasses contg Cu and Fe scarcely seem to be affected by the Ra ray. Ni and Co glasses, on the contrary, are changed in a very perceptible manner. C. W. OWINGS

The making of mirrors by the deposition of metal on glass. ANON *Bur Stand-ards Circ* No 389, 17 pp (1931) —This publication supersedes *Jetter Circ* No 32 and deals principally with the methods for producing mirrors by the chem. decompn of Ag on glass. The Brashear, Rochelle salt, and formaldehyde formulas are given, also a detailed discussion of the precautions necessary to avoid danger, and the technic necessary to obtain the most satisfactory results. Reflecting films are produced on glass by the chem. decompn of Cu, Pt or PbS, by cathode sputtering and by condensation of vaporized metals. W. H. BOYNTON

Preparation of mirrors by sputtering metals on to glass surfaces. A C G BEACH *J Sci Instruments* 7, 193-5(1930) —Exptl details are given for the prepn of mirrors of Ag and Ni by sputtering. B C. A.

Specifications for tank blocks. W. J REES *J. Soc Glass Tech* 14, 124-30P (1930) See C. A. 25, 1051. H. F. K.

The corrosion of tank blocks by opal glasses. C J UETTMANN AND S. M SLATER *J Am Ceram Soc* 13, 931-4(1930) —More opacifier must be used in batches for day tanks than for closed pots. The glass used was: SiO_2 65, B_2O_3 1.5, Al_2O_3 5.5, CaO 11.5, K_2O 6.5 and Na_2O 10.0%. During a tank failure a fluid of lower viscosity ran out before the glass came through. A similar vitreous substance was found between the glass and the bottom block. Apparently certain of the corrosive ingredients melt out first and percolate through the batch to the bottom. Cooling the blocks and filling in 3 sep charges decreased block stones and increased tank life. C H KEAR

Mechanical handling of raw and manufactured material under the conditions obtained in glass works. T. E DUNLEAVY *J Soc Glass Tech* 14, 203-75T(1930) H. F. K.

Some experiments on the cascade method of melting glass. T C MOONSHHEAD AND E. A COAD-PRIOR *J Soc Glass Tech* 14, 249-60T(1930) H. F. K.

Efficiency of tank furnaces for glass melting. Proposals for the standardization of its nomenclature and methods for its measurement. W FRIEDMANN *J Soc. Glass Tech* 14, 91-102P(1930) H. F. K.

Notes on the effect of load on the fuel consumption of a glass tank furnace. E A COAD-PRIOR *J Soc Glass Tech* 14, 84-90P(1930) —For each 10% increase in the load over a range of 10-100% the increase in the fuel consumption was about 1.9%. H. F. K.

Effect of factory organization and size of pots on the production from pot furnaces. W W WARREN *J Soc Glass Tech* 14 103-6P(1930) H. F. K.

Manufacture of working life of glasshouse pots. O BARTSCH *J Soc Glass Tech* 14, 107-18P(1930) —A comparison of pots contg the siliceous clays like the Grossalmerode clay and those of higher Al_2O_3 content like the Rakonitz clay showed the siliceous bodies to have higher expansion up to 800°, lower contraction after 1400°, less resistance to thermal shock at temps below 600°, a higher softening temp but a lower temp. of complete collapse, and a lower resistance to glass corrosion, though the glass produced was of better quality than that in the more aluminous pots. H. F. K.

Manufacture of glasshouse pots in England. G V EVERA *J Soc Glass Tech* 14, 119-23P(1930) —Very good results are obtained with the Old Mine Stourbridge fireclay which contains up to 72% SiO_2 (ignited basis) and about 25% grog of the same material. H. F. K.

Contribution of technical education to ceramic progress. EDWARD ORTON, Jr *Trans Ceram Soc* 29, 415-36(1930) H. F. K.

Color changes in pastes. G MILANI *Correre ceram* 11, 453-7(1930) —The color changes are generally due to Fe, Ti, and V. Under oxidizing conditions the color caused by the Fe may vary from yellow to brown, depending on the amount of Ca present. S in the fuel reduces the decoloring power of the Ca, the S combines with the Ca with the formation of white and yellowish colored spots. Sol salts tend to conc. on the surface during drying, where they cause stains of insol. sulfates by taking up the S from the gases of combustion. R D BUMBACHER

Shaft furnaces for calcining clays and silica. ALBERT HIRT *Ceramique* 33, 272-80(1930) A J MONACK

Thermal expansion of silica bricks. BERNARD LONG *Ceramique* 33, 268-72 (1930) —The results are expressed graphically. A J MONACK

The influence of atmosphere on the load-bearing capacities of fire bricks. S C 11 SWALLOW *Refractories J* 6, 83-94(1930) —Load tests were made on specimens from 5 types of fire bricks. All samples showed increased deformation rates in a reducing atm. The effect was greatest on samples with a high Fe content. A change from re-

ducing back to oxidizing conditions restored the original rate of deformation. The effect of steam was negligible. M. A. EODY

Comparative effects of glass batch, cullet, soda and sulfate on fireclay bricks. P. P. БУДНИКОВ. *Feuerfest* 5, 181-3(1929).—Ten bricks of Russian manuf. of the type used for the bottom and sides of glass tanks were studied. The chem. compn., the refractoriness under load, and the porosity were detd., the chem. resistance was studied on small cubes in which circular cavities had been bored. The cavities were filled with the reagents, and the specimens were then heated in an elec. furnace and held at 1430° for 2 hrs. The glass-batch mixt. was slightly more active than the cullet, but in neither case was corrosion very marked. The sulfate showed the most corrosive action, especially in the presence of C, since the Na-S formed reacted vigorously with the refractory material. Soda also proved a powerful corrosive. In every case the degree of corrosion was dependent on the porosity of the material. B. C. A.

Progress in klinker (brick) construction in South Germany. A. LANG. *Ziegelwelt* 61, 511-4(1930) H. F. K.

German klinker. W. ANGUS McINTYRE. *Trans. Ceram. Soc.* 29, 472-6(1930).—German plant processes in the manuf. of "Klinker" are described. The term covers a no. of hard burned, all-clay bricks, with a variety of colors, produced by maintaining alternately oxidizing and reducing atms. in the kiln. H. F. K.

Factors in the manufacture of porcelain. MARC LARCHEVÈQUE. *Ceramique* 33, 221-4(1930) A. J. MONACK

Manufacture of roof tile. MILDE. *Ziegelwelt* 61, 545-8(1930) H. F. K.
Roof tile continuous drier with a horizontal alternating draft. E. GAOSZKINSKY. *Ziegelwelt* 61, 560-1(1930) H. F. K.

Artificial kaolin. JAROSLAV MELBAUER. *Chem. Obsor* 5, 236(1930).—From chem. analysis, checked by x ray spectrographic tests, the so-called "artificial kaolin" was found to be a waste from the chem. treatment of bauxites by alk. process. It contains only 20% SiO_2 and is apparently an amorphous mixt. of $\text{Al}(\text{OH})_3$ and SiO_2 . J. K.

Influence of high temperature on Prosyannaya kaolin. V. V. YURGANOV AND M. V. ZUSMANOVICH. *Trans. Ceramic Research Inst. (Moscow)* 1929, No. 21, 28-55(in German 56-7).—At a burning temp. of 850° reagents cause no changes in kaolin differing from those caused in similar materials. At this temp. kaolinic acid anhydride (metakaolin) apparently exists. At 950° kaolinite decomposes into free SiO_2 and the difficultly sol. form of clay and shows the first sign of formation of one or more Al silicates. At 1050-1100° there is little change except further combination of free SiO_2 and Al_2O_3 . No sillimanite was found. At 1200° the difficultly sol. residue increases; it has the compn. $5\text{Al}_2\text{O}_3 \cdot 4\text{SiO}_2$. At 1320° the ratio is $\text{Al}_2\text{O}_3 \cdot \text{SiO}_2 = 2.70$ 2, or $4\text{Al}_2\text{O}_3 \cdot 3\text{SiO}_2$. Kaolin burned at 1400° shows some crystals when examd. in thin layers, increasing at 1470°. At 1400° and above, $\text{Al}_2\text{O}_3 \cdot \text{SiO}_2 = 3 \cdot 2$ (approx.). E. M. SYMMES

Ceramic investigations of Tshasoff-Jar clays. K. KÖHLER. *Trans. Ceramic Research Inst. (Moscow)* 1929, No. 20, 88-150(in German 187-92).—Great care was taken to obtain a representative sample from agitated slime. Improper expulsion of free and combined water is discussed, and examples and tables are given. Plasticity is detd. by a method based on the Brongniart principle. Tensile strength is shown in tables. Bonding value is detd. by a special app. Clays are grouped according to rate of drying. Total shrinkage was detd. at 400-1410° and plotted in curves. Porosity was detd. on samples burned at the above temps. and shown in curves. The character of porosity was detd. by microstructure, sp. gr. and wt. vs. vol. Up to 900° the samples have the same porosity as before burning; above 900° secondary pores form, increasing in no. and size with temp. rise. At 900° there are mainly open pores, but above this temp. there is sintering. At 1200° the clay is completely sintered, the secondary pores closing. Above this temp. the gas trapped in the pores enlarges them. Refractoriness was detd. by comparison with Seger cones. A rational classification of clays is proposed. E. M. SYMMES

Use in the ceramic industry of the sandy Tshasoff-Jar clay called "Balyki." P. N. KAVOKIN. *Trans. Ceramic Research Inst. (Moscow)* 1929, No. 20, 172-81(in German 194-5).—"Balyki" is a sandy clay, removed as overburden and discarded hitherto. Two samples fired at 1250° showed a total shrinkage of 7-10%, water absorption 1.12%, porosity up to 2.3%. Such clays can be used to make stoneware. About 60-65% of fireclay can be added to them, then fired at 1200° to make fire brick having 4% total shrinkage. Repeated firing causes no after-shrinkage. Resistance to abrupt temp. change is satisfactory. Water absorption is only 5%, and tensile strength 400-500 kg./sq. cm. A table shows deformation under load. Such bricks will not withstand furnace temp. above 1200°. E. M. SYMMES

Feldspar. I. J. FAIRCCHILD, GEORGE K. BITTGEN AND R. P. LAMONT. *Bur Standards Commercial Standard C S 23-30*, 15 pp (1930).—Classification is based on particle size and chem. compn. Chem. classification is in 3 groups: (1) includes ceramic grades based on percentage of SiO_2 and alkali ratio, and contg less than 4% Na_2O , (2) includes spars used for glazing, based on percentage of Na_2O , and contg $\geq 4\%$, or more Na_2O , (3) includes spars used for glass making based on SiO_2 , Al_2O_3 , and R_2O . Tables are given with grades and specifications. Standard methods of phys. and chem. tests are given in detail. Chem. methods include analysis for SiO_2 , " R_2O ," Al_2O_3 , Fe_2O_3 , CaO , MgO , K_2O , Na_2O and ignition loss. Acceptors of this standard are listed. The effective date for production under this classification is Sept. 1, 1930.

ALICE W. EPPESON

Ceramic bodies described according to their mineral compositions and as silicates. JOSEF HOLZ. *Sprechsaal* 63, 929-31 (1930).—A supplement to similar work by W. appearing in C. A. 22, 852.

C. H. LORIG

Refractory materials for electric furnaces. I. The melting of lead and its alloys. ALFRED B. SEARLE. *Metal Ind* (London) 38, 3-4 (1931).—The max temp attained in an elec. furnace when melting Pb or Pb alloys is about 1000°, and generally much lower. Elec. furnaces are the most suitable for melting this class of alloys. The properties of Elec. furnaces are outlined. Magnesia brick is suitable for parts of a Pb furnace in contact with slag. For most induction furnaces dead burned magnesia is best with pitch or colloidal MgO as binding material. A melted mixt. of PbO and clay applied as a paste affords protection to the refractories in high frequency furnaces for these fusible alloys. Silica bricks are used in the roofs. Most refractory materials are good conductors of electricity at red heat, this often results in serious leakage of current and may cause a serious defect in the furnace.

W. H. BOYNTON

The vitreous enameling of cast iron. J. H. D. BRADSHAW. *Foundry Trade J.* 44, 123-4 (1931), cf C. A. 24, 2261.—A review.

E. H.

The microscopic structure investigation in service in the enamel technic. FRITZ KRAEPE AND VIKTOR LINDT. *Sprechsaal* 64, 1-4 (1931).—Metallographic methods show defects appearing between the ground coat, cover coat and the iron.

C. H. L.

Use of nepheline for enamels. A. I. ZIMLIN. *Keram. i. Stklo* 6, 495-8 (1930).—Tests showed that nepheline from the Ural mountains can be used with success instead of feldspar in ground or white enamels.

M. V. KONDOIBY

The effect of furnace gases on the quality of enamels for sheet steel. ANDREW I. ANDREWS AND EMANUEL A. HERTZELL. Univ. Ill. Eng. Expt. Sta., *Bull.* No. 214, 29 pp (1930); cf C. A. 24, 5957.—Small quantities of SO_2 such as are sometimes found in the atm. of industrial districts may be deleterious in enamel baking through its etching or pock marking action.

H. L. OLIN

Ashes and the remains of fuel found by the excavation of ruins of a seventeenth century glass factory near the Keizergracht at Amsterdam (Keeulen) 2.

Taschenbuch für Keramiker, 1931. Band I. Textband. FRITZ H. ZSCHACKE. Glas. H. J. KARMAUS. Email. 248 pp. Band II. Alph. Führer durch der Keramik, Glas- und Emailindustrie mit Bezugsquellen. 178 pp. Berlin: Verlag Keramische Rundschau. Linen, M. 4.

Glass making. SOC. ANON. DES MANUFACTURES DES GLACES ET PRODUITS CHIM. DE SAINT GOBAIN, CHAUNY & CIÉ. Fr. 692,932, June 28, 1929. Means for controlling the temp. of app. used for rolling glass.

Apparatus for feeding mold charges of molten glass. EVERETT S. GREER (to Hazel Atlas Glass Co.). U. S. 1,790,984, Feb. 3. Structural features.

Glass tank. KARL G. KUTCHKA. U. S. 1,790,820, Feb. 3.

Anchoring device for use in drawing sheet glass. ARTHUR E. SPINASSE. U. S. 1,790,774, Feb. 3. Structural features.

Apparatus for making sheets of glass. MISSISSIPPI GLASS CO. Fr. 693,889, April 14, 1930.

Apparatus for rolling glass. CLAUDE LOCREILLE. Fr. 692,500, Mar. 21, 1930.

Apparatus for annealing glass. SOC. ANON. DES MANUFACTURES DES GLACES ET PRODUITS CHIM. DE SAINT GOBAIN, CHAUNY & CIÉ. Fr. 692,971, July 5, 1929.

Treating frosted glass articles. ROWLAND D. SMITH (to Corning Glass Works). U. S. 1,791,066, Feb. 3. Articles such as lamp bulbs are washed in an aq. soln. of a simple Al salt of an aliphatic org. acid such as Al tartrate in order to prevent or lessen weakening. Cf C. A. 24, 5124.

Wire glass production. GEORGE AURIEU (to Mississippi Glass Co.). U. S. 1,791,260, Feb 3 Mech. features

Apparatus for treating sheets of reinforced glass. DUPLATE CORP. Fr 693,395, April 5, 1930

Laminated glass. JAMES F WALSH and JOHN H STEVENS (to Celluloid Corp.) Can. 308,522, Feb 10 1931 Sheets of glass have interposed therein a pyroxylin sheet contg cellulose nitrate having a N content of 10.5-11.5% 3-50 parts of camphor to 100 parts of cellulose nitrate, urea, a residual solvent comprising volatile alc and a coloring matter opposed to yellowness

Safety glass. (Miss) M SCHMIDT. Belg 371,936, Aug 31, 1930 One or more sheets of cellulosic material are placed between two or more sheets of glass without interposition of adhesive The combined sheets are immersed into one or more baths which transform the surface of the cellulosic material into a film which can adhere to glass The baths are composed of glycol monochlorohydrin, BuOAc , etc

"Splinterless" glass sheets. JOHN JEFFRAY. Fr 693,127, April 1, 1930 See Brit. 327,949 (C A 24, 5124)

Mixing clay and chamotte. W LOTHES NACHFOLGER MAX STEPHAN. Ger 515,171, Jan 18, 1925 The chamotte is sprayed with water in a screw conveyor which delivers it to a wet mixer The clay is fed to the mixer by another conveyor

Bricks. ALFRED PAUL, JR. U. S. 1,791,372, Feb 3 A highly porous light material such as volcanic ash is broken into small particles and 85-90% of this material is mixed with 15-10% of unslaked lime and water is added and the materials are agitated together in a closed space in which the steam generated is retained under pressure the material is molded into bricks and the latter are subjected to the action of steam under pressure.

Piling brick for burning. ARTHUR J. THIERRIEN (to National Brick Co.) U. S. 1,790,882, Feb. 3. Mech. and structural features

Tile and similar products. DAVID A. CABLE (to U. S. Quarry Tile Co.) U. S. 1,791,234, Feb 3. A block of moist material is dried to beyond the plastic limit, trued and then burned. An arrangement of app. is described.

Ceramic vessels. MAX HAUSER. U. S. 1,790,918, Feb 3 Ceramic vessels of good toughness and heat cond. are made by shaping and firing a mixt. of a ceramic raw material and a Cr alloy such as ferro-Cr and Si

Decorating ceramic objects. DEUTSCHE GOLD- UND SILBER-SCHNEIDANSTALT VORM ROESSLER. Fr. 692,562, Mar. 22, 1930 Ceramic objects are decorated with Au to which Ag, Cu or Bi may be added, the baking of the powder being effected in the presence of metals capable of increasing the resistance of the Au to the fire, such as Rh, Cr, U and like metals Cf C A 24, 215

Decorating pottery. DEUTSCHE GOLD- UND SILBER-SCHNEIDANSTALT VORM. ROESSLER. Brit. 335,788, Nov. 9, 1928. The process of decorating described in Brit. 308,226 (C. A. 24, 215) is modified by the use of ceramic enamel colors, and thus veins of a darker shade than the groundwork, or of variegated shades, can be produced. Various details and formulas of coloring compns are given Brit 335,789 relates to the use of similar processes for coloring metal surfaces such as Al

Apparatus for the continuous baking of ceramic products. FOURMENT ET LADURÉE (S A s. L.). Fr. 692,643, June 22, 1929

Abrasive. BAKELITE CORP. Fr. 692,733, Mar. 25, 1930 An abrasive is made by mixing abrasive grains, a finely divided substance and a resinoid capable of being hardened The mixt. is molded and heated until the resinoid hardens Cf C A 24, 2569

Abrasives. NOBSON CO. Fr. 692,738, Mar. 25, 1930. An abrasive consists of a cryst α -alumina free from non-aluminous impurities and β alumina Cf C A 24, 2854.

20—CEMENT AND OTHER BUILDING MATERIALS

J. C. WITT

Studies on calcium ferrites and iron cements. IV. SHOICHIRO NAGAI AND KATSUHIKO ASAKA. *J. Soc. Chem. Ind., Japan* 33, Suppl. binding, 256-7(1930); cf. C. A. 24, 4605.—Analyses of mixts. of the 3-component system, $\text{SiO}_2\text{-Fe}_2\text{O}_3\text{-CaO}$, after heating at increasing temps. showed that Ca ferrite formed first, and later SiO_2 and CaO combined to CaO SiO_2 and then to 2CaO SiO_2 and 3CaO SiO_2 . V. *Ibid* 312-5.—The combination in Fe cements is hardly completed at 1450-1500°, which is higher than the sintering temps. of ordinary portland cement mixts. and nearly 100-

150° higher than those of Kuehlzement raw mixt. Compression tests on the synthetic mixts used in these expts show somewhat smaller strengths than those of ordinary and special portland cements owing to the large contents of 2CaO Fe_2O_3 and small contents of 3CaO SiO_2 .
H. G. PARISH

Examination of road material with small samples. R. GÄRNGG *Mikrochemie* [N S] 2, 281-92(1930) —At the Technische Hochschule at Vienna it has sometimes been desirable to test small samples of materials used in road construction or of the pavement itself. In this paper some interesting information is given concerning grain size after pulverizing, and 2 instruments are shown: a ram for prep and testing small samples and a machine for testing small samples under pressure and under bending strain.
W. T. 11

Report of Committee XVII on wood preservation. F. C. SHEPHERD, et al. *Proc Am Ry Eng Assoc* 1930, 678-770 —Remission of manual —Tables prep'd by the Bureau of Standards for correcting the vol and sp gr of creosote oils for different temps are presented for adoption as standard. Definitions used in wood preservation: A final revision of the list of important terms in the industry is presented. Service records —Data of completed service tests of ties, supplementing previous tables, are given, the annual table for tie renewals has been extended to include 27 railroads to 1928, reports on special test tracks on 7 large trunk lines are submitted. Piling for marine construction —Progress reports of exposure tests of treated and untreated wood specimens for resistance against marine borers are submitted. Creosote-petroleum treatments —Inspection of creosote petroleum and zinc petroleum ties on the Atchafalaya, Topeka and Santa Fé R R is reported. Termite destruction —An extensive report is given, with comprehensive bibliography, on termite injury to wood in railway structures and on methods of prevention. Loss of preservative in treated ties due to oil burning weed destroyers —Preliminary tests indicate very small loss from ordinary operation of such machines, no final conclusions are drawn.
ALFRED L. KAMMERER

Report on work on protection of timber during the year 1929-30. GEORGE BARGER Dept Sci Ind Research, 11th (interim) Rept Comm Inst Civil Eng. Deterioration of Structures in Sea Water, 1930, 13-5, cf. C. A. 24, 2876 —Several of the tests of treated timbers exposed to marine borer attack at Colombo and Singapore have been completed, the remainder being still under observation. The conclusions to date are: Creosote to which D. M. (chlorodihydrophenarsazine) has been added shows no advantage over creosote alone. With the EtOH solns of 4 arsenical compds 2% (and in most cases 0.6%) of D. M. gives protection in the treated portions of the wood, with less active poisons and lower D. M. contents there is serious attack; fuel oil as a vehicle yields unsatisfactory penetration. Pintsch gas tar gives excellent penetration and is suggested as a vehicle for toxics.
ALFRED L. KAMMERER

Impregnation of the ninth series of timber test pieces for exposure at Kilindini and Mauritius and experiments on the impregnation of incised timber. S. M. DIXON Dept Sci Ind Research 11th (interim) Rept Comm Inst Civil Eng. Deterioration of Structures in Sea Water, 1930, 16-8, cf. C. A. 24, 2876 —Swedish fir blocks were treated with naphthalene alone and in various combinations with creosote by the Bethell process and exposed to borer attack at Kilindini and Mauritius. Tests with incising Oregon fir timbers to a depth of $\frac{3}{4}$ in in staggered rows 15 in apart center to center showed a very considerable increase in penetration of creosote with little reduction in strength.
ALFRED L. KAMMERER

Termites and termite damage with preliminary recommendations for prevention and control. S. F. LIGHT, MERLE RANDALL AND FRANK G. WHITE. Calif Agr Expt Sta., Circ 318, 3-61(1930) —Coal tar creosote or ZnCl_2 used thoroughly to impregnate wood will protect lumber, posts, ties etc., effectively against termites. Cyanide fumigation is not a satisfactory control measure. *a*-Dichlorobenzene is fairly effective and does not constitute a fire hazard. When a 10% soln of Na_2HAsO_4 is liberally used on soil or in masonry foundations, termites are effectively controlled. A nontoxic ground treatment consists of a 10% soln of CuSO_4 or ZnCl_2 .
C. R. FELLERS

Determination of Mg in portland cement (REDMOND, BRIGHT) 7. Natural weathering and a comparison of chemical and natural weathering of building stones (KAISER) 8. Improved asphalt extractor (STANTON) 22. Ktn for preheating and clinkering of cement material (U. S. pat. 1,791,165) 18. Preserving wood, glue, etc. (Brit pat. 335,627) 15.

Cement and Concrete; A General Reference Book. Chicago Portland Cement Assoc. 64 pp

KNUCHEL, HERMANN Untersuchungen über den Einfluss der Fällzeit auf die Eigenschaften des Fichten- und Tannenholzes. *Tl. I. Der Einfluss der Fällzeit auf einige physik. und gewerbl. Eigenschaften der Holzes.* Bern Böhler 127 pp F 5

Cement. ARTHUR ANKER Fr 693,552, April 8, 1930 The crude materials used for making cement by heating to the commencement of fusion are stored for a certain time, with or without exposure to steam, in ripening silos heated or insulated from the cold

Portland cement. THE ASSOCIATED PORTLAND CEMENT MANUFACTURERS, LTD Fr 693,200, Apr 2 1930 A white portland cement is obtained from colored raw materials by treating the cement during or immediately after the calcination and before the cement has cooled with a reducing agent, after which the cement is rapidly cooled to a temp preventing fresh oxidation The reduction may be carried out by projecting finely divided coal on to the cement as it leaves the furnace or by emptying the red hot cement on to a surface covered with a film of paraffin

Manufacture of artificial portland cement and similar products. R FOURDAX Belg 371,579, Aug 31, 1930 The paste or slurry to be burned is introduced at or near the top of the stack through which the gases from the kiln are discharged, so as to recover the heat of the gases and prep the paste for the burning Grates or perforated plates are provided in the stack to distribute the material uniformly

High-alumina cement from phosphate rock and alundite. HERBERT H MEYERS (to Armour Fertilizer Works) U S 1,791,103, Feb 3 A mixt of finely ground phosphate rock and finely ground alundite is calcined, and the calcined product is treated with NH_3 water and CO_2 , the solid residue is sepd from the soln, and the sepd residue is burned to produce a cement of the high alumina quick-hardening type Cf. C A 24, 5905

Color composition for cementitious products. HERBERT E. BEDWELL Can 308,454, Feb 10, 1931 A coloring compn for cementitious material comprises 30 lb carbon black, 4 lb Fe oxide and 100 lb of water to form a paste capable of retaining the color suspended evenly throughout the mass. One lb of oil of pine is preferably added

Process for the regeneration of stones, marbles and concrete products. F. RICHER Belg 371,673, Aug 31, 1930. The material is impregnated with a metal fluosilicate contg a slight excess of acid.

Apparatus for alaking the excess of lime in cement or clinker. ARTURO MALIGNANI and CAMILLO MALIGNANI Ger 514,474 April 3, 1928

Rotary kiln for burning cement, ores, etc. LIONEL D. PARKER (to Vickers, Ltd) U S 1,791,282, Feb 3

Concrete. VLADISLAV DARYNK Fr 693,446, April 7, 1930 A silicate concrete is made by mixing a powd basic silicate rock with a sol glass and sand to make a mortar which is mixed with water and ground, and then broken stones or gravel are added

Concrete grindstones for wood-pulp manufacture. F RUELHMAN Brit 335,521, June 21, 1929 The external temp surrounding the stone is made equal to the internal temp of the stone during its manuf in order to avoid unequal stresses developing during the setting and hardening Various details of manuf. are described

Bituminous materials. ALUMINUMWERK BERGHAU UND INDUSTRIE A-G. Fr 693,867, Apr 14 1930 Bituminous materials, particularly for roads, are made by incorporating finely divided bauxite with bitumen, etc Other substances such as clay, schist, limestone or sand may also be added

Road-surfacing material. SOUTH METROPOLITAN GAS CO., H. PICKARD and H. STANIER Brit 335,668, July 23, 1929 Coal, 1-10 lb., is dispersed in tar, 100 lb., (which may contain up to 10% of added pitch) at 250-300°, and 30-60% of dried and warm finely divided mineral matter is added and admixed

Road-building aggregate. LESTER WITTENBERG (to Barrett Co) U. S. 1,791,109 Feb 3 Stone and bitumen are mixed while the bitumen is fluid and the stone has a temp approx equal to that at which the bitumen will set An arrangement of app is described

Tar macadam. J. F. WAKE Brit. 335,525, April 24, 1929. Aggregate is first coated with bituminous material and then with tar or a tar compn comprising 20-75% of the total material in the 2 coatings, and a third coating of oil or low-viscosity tar also may be applied The tar material used contains about 0.5% tar acids and 2% of naphthalene

Roofing. G. F. JAULET and O. PARDOEN Belg 372,317, Aug 31, 1930 Asbestos

is mixed and worked up with a vegetable oil together with a binder for the oxidizing substances, so that the asbestos fibers will be enveloped in completely hardened and oxidized oil

Material suitable for shingles formed of cement and fibrous material. **JOHN W. LEDBETTER** (to Amherst Asbestos Shingle & Sheathing Co.). U. S. 1,700,822, Feb 3 A material which may be formed mainly of cement and fibrous material is provided on its surface with a plurality of variably overlapping streaks comprising powdery coloring matter of different hues and cement irregularly distributed

Drying stand for wood. **OLOF ENGELBREKTSON** Ger 515,450, Jan 7, 1928.

21—FUELS, GAS, TAR AND COKE

A. C. FIELDNER AND ALDEN H. EMERY

Fuel utilization in 1930. Fuel Research Inst *J Fuel Soc Japan* 10, 4-36(1931). English abstr. 1-6—A review **F. I. NAKAMURA**

The utilization of Brazilian fuel. **FRANCIS W. FAIRBANKS** *Chaleur et ind* 11, 320-6(1930)—The utilization of coal, coal washing and briquetting, powdered coal, carbonization, manuf. of metallurgical coke, gas manuf., and low temp. carbonization in Brazil are generally treated. Lignites, bituminous schists and peat in Brazil are briefly mentioned. Liquid and gaseous fuels are also briefly treated. A bibliography of 15 references is given **S. L. B. FRIERSTON**

Fuels and their combustion. **ROGER MARTIN** *Chaleur et ind* 10, 149, 206-12(1929) **S. L. B. FRIERSTON**

Some factors affecting the problem of smoke prevention. **W. E. GIBBS** *J Inst Fuel* 3, 361-8(1930)—A theoretical discussion of the formation of smoke and its elimination by (1) gravity sepn., (2) pptn. by centrifugal force, (3) washing and (4) electrostatic sepn. **D. A. REYNOLDS**

The relative merits of coal and oil in the mercantile marine. **W. M. MEYER** *J Inst Fuel* 4, 54-6(1930)—A discussion **D. A. REYNOLDS**

Fuel problems in the mercantile marine. **STESSY D. FREEMAN** *J Inst Fuel* 4, 49-53(1930) **D. A. REYNOLDS**

Fuel problems in the mercantile marine. **W. J. MULLER** *J Inst Fuel* 4, 59-63(1930) **D. A. REYNOLDS**

Diesel-engine developments. **MAXIMILIAN GRACKER** *J Inst Fuel* 4, 16-20(1930) **D. A. REYNOLDS**

Two important physical properties of fuels pertaining to the formation of mixtures in carburetor engines. **WAWRZONOK** *Automobiltech Z* 33, 202-4, 317-8, 364-6, 408-10, 478-80(1930)—The suitability of fuel such as gasoline, benzene, alc., etc., and their mixts., for carburetor engines is detd. by the "evaporation no." which is obtained by exposing the fuel to a temp. of 45° and a pressure of 80 mm., adding the time (in min.) required to evaporate 5, 10, 15, 20, 25-95% of fuel and dividing by the number of time readings, 4 e., by 19. An app. constructed for this detn. by the author is described in detail **A. A. BOWLING**

The utilization of gas oil in motor trucks. **A. GAEDEL** *Chaleur et ind* 11, 277-85, 323-43(1930)—G. gives an account of Chelowsky's app. for oil gas. This gas contains CO₂ 4.8, CO 13.8, H₂ 8.6, CH₄ 9.1, C₂H₆ 4.4 and N₂ 59.3%, and the thermal yield is 70%. The gas is fed to a four-cylinder motor and its performance is analyzed. The prepn. of gas oil by "the chemical carburetor" is described with sketches and data, and a general summary is added **S. L. B. FRIERSTON**

Methane as motor fuel. **J. BAERN** *Brennstoff-Chem* 12, 27-9, 45-7(1931)—Uses for the large amts. of coke-oven gas which have been freed of H₂, etc., for various industrial purposes are discussed from practical and economic viewpoints. Direct use as fuel for internal-combustion engines is proposed, previous attempts thus to employ gaseous fuels are reviewed. The combustion of CH₄ and homologs is examined, and analytical data for the exhaust gases from fuel contg. 80% CH₄ are given. Several months' road test indicates that 0.85 cu. m. CH₄ = 1.0 l. C₂H₆, that the motor is not injured and that less lubricating oil is required. Cost data are given for the use of various types of steel cylinders and high pressure piping from plants to population centers **F. W. JUNO**

Recent experiments on the pyrolysis of methane. **R. V. WHEELER AND W. L. WORN** *Fuel in Science & Practice* 9, 567-74(1930)—Addnl. exptl. work on the pyrolysis of CH₄ is described (cf. C. A. 24, 5715). Two types of expts. were made, (1) in which

C_{12}H_2 was heated in a quartz bulb of 60 cc capacity at different temps., pressure changes being observed by manometer readings and (2) in which large vols. of C_{12}H_2 were passed through heated tubes of different materials, the products of decompn. being collected and analyzed. The rapid initial decompn. of C_{12}H_2 is chiefly a surface effect, but the decompn. does not proceed to an equil. state by surface reaction because of the formation of a protective layer of adsorbed H on the silica surface, which leads to a false state of equil. Fe has no effect on the decompn. of C_{12}H_2 , but at higher temps. it has a marked catalytic effect on the pyrolysis. Tars and C_2H_2 are produced during C_{12}H_2 decompn., hence the reaction is complex. When C_{12}H_2 is decompd. by prolonged heating, C and H are the principal products. As the duration of heating is decreased, more of the decompd. C_{12}H_2 is converted into higher hydrocarbons. For periods of heating of 0.3 sec. at 1050° in a silica tube, 84% of the decompd. C_{12}H_2 produced higher hydrocarbons. The shortest periods of heating favor the formation of mols. of 2 C atoms, whereas on increase in the duration of heating these diminish, and aromatic compds. are formed. C_{12}H_2 and C_2H_2 were not detected among the decompn. products. C_{12}H_2 flowing through an Fe tube was completely decompd. into its elements at comparatively low temps., yielding a C deposit quite different from that deposited in tubes of SiO_2 or chrome-Fe.

D. A. REYNOLDS

Liquefaction of methane according to recent investigations. JOSEF DUPUIS. *Przemysl Chem.* 14, 302-9 (1930).—This is a review dealing principally with liquefaction of methane and production of motor benzines. It contains numerous patent and periodical literature references.

A. C. Z.

Purification of acetylene. S. KAKUTANI AND A. YAMADA. *Repts. Imp. Ind. Research Inst., Osaka, Japan* 11, No. 11, 78 pp. (1930).—A review on the method of removing impurities such as PH_3 , H_2S and NH_3 in acetylene and some exptl. data are given.

F. I. NAKAMURA

Calculation of the heating value of solid fuels. F. SCHUSTER. *Glückauf* 67, 232-5 (1931).—Proposed formulas, based on proximate and ultimate analysis and on air required for combustion, are reviewed and tested on typical coals, the detd. heating values being known. Formulas based on air required and on the proximate analysis give only roughly approx. results. Vondracek's formula, $\text{H}_0 = 81\text{C} + 342\text{H} + 22.5\text{S} - 36.60\text{Cal./kg.}$, and that of S, $\text{H}_0 = (254 + 0.3550)[\text{C}/3 + \text{H} - (\text{O}-\text{S})/8]\text{Cal./kg.}$, gave results within 2% of the detd. value. The tests covered bituminous coals, brown coals, peat and wood.

J. D. DAVIS

Softening of ash of solid fuel. H. A. J. PIETERS. *Chem. Weekblad* 27, 331-1 (1930).—A modified app. for detg. the softening point of ash is described, and a large no. of results are recorded. Three classes are distinguished, with softening temps., resp., below 1250° , and above 1250° , the 3rd class showing softening but not melting above 1250° . The relationship between softening temp. and chem. compn. is discussed.

B. C. A.

Powdered fuel, with special reference to Lancashire boilers. H. J. HILL. *Trans. Inst. Mining Eng. (London)* 80, 158-79 (1930-31).—The max. fineness possible is the aim in prep. pulverized fuel; however, 75-85% through 200-mesh is the usual size used. Air-float anthracite dust burned more freely than dust from a pulverizer. A microscopic examn. showed that the dust from the pulverizer was rounded whereas the float dust was angular and hence exposed more surface during combustion. Present-day burners for boilers and furnaces are designed to obtain max. turbulence and in most cases the shortest possible length of flame. The operation of 2 Lancashire boilers has been entirely satisfactory.

C. W. OWINGS

Pulverized fuel for the small unit-shell type boiler, metallurgical and chemical processes. H. W. HOLLANDS AND E. C. LOWNDEN. *J. Inst. Fuel* 3, 225-30 (1930).—Pulverized coal is compared to oil and gas as a fuel in cement plants, road tar plants and for evapn. in the manuf. of NaOH .

D. A. REYNOLDS

Tests of a Ljungström air preheater with a chain grate stoker. W. SCHULTES. *Glückauf* 66, 11-7 (1930).—Comparative tests were made on a Babcock and Wilcox water tube boiler of 122 sq. m. heating surface with and without the air preheater (regenerative) which was located at the rear of the boiler in the flue-gas canal. The fuel used was low grade, being a 1:1 mixt. of fine coal and coal of a medium grade. The mixt. gave 14-27% ash, about 14% H_2O and 5400 to 5700 cal. Mech. imperfections of the preheater caused considerable trouble. The main advantages were increased capacity and easy kindling of the low-grade fuel. It did not enhance clinker trouble. The operation of the preheater was not interfered with by dust (ash) in the flue gas. Complete test data are given.

J. D. DAVIS

Low-temperature carbonization and the production of electricity in Germany.

P. ROSIN *J Inst Fuel* 3, 189-98(1930), cf *C A* 24, 223.—A review is given of the present trend in Germany toward the coupling of carbonization plant, power station and gas works. D. A. REYNOLDS

The x-ray stereoscopic examination of coal. ARTHUR N. WILSON *J Inst Fuel* 3, 433, 4, 64-5(1930), cf *C A* 24, 5969 D. A. REYNOLDS

The history and present position of coal investigation by oxidation. W. FUCHS. *Fuel in Science & Practice* 9, 581-6(1930)—A review D. A. REYNOLDS

Rational sampling of coal for chemical analysis. K. KLING AND J. PFANHAUSER. *Przemysl Chem.* 15, 12-6(1931)—Directions and care to be exercised in taking samples of coal from geologic deposits, cars, classifiers, storage, etc., are given, and the method of prep. and cutting down the large sample is described in detail. A. C. ZACHRY

Apparatus for sampling powdered coal. D. J. W. KREULEN. *Chem Weekblad* 28, 66(1931)—Sampling of a carload of powd. coal with a shovel is inaccurate. A sampler is described made up from a can. A small hole in the bottom makes it possible to see if the can is full. E. SCHOTTE

Constant factors for the calculation of the calorific value of Cape Breton coals from proximate analysis data. J. L. BOWLEY *Trans Can Inst Mining Met* 32, 489-90 (1929)—A comparison of various formulas. The expression preferred is $1x + 145.5C + 40.5S = B$ t u per lb., where V is the volatile matter, C the fixed C and S the S per centage, and x is an arbitrary const. depending on the colliery. B. C. A.

Apparatus for determining the tendency of coal and other materials toward spontaneous combustion. D. J. W. KREULEN. *Chem Weekblad* 27, 690-4(1930)—See *C A* 25, 394 E. SCHOTTE

Determination of the expansion pressure of coal. H. NIDELMANN. *Brennstoff-Chem.* 12, 42-3(1931)—The app. of Baum and Heuser (cf *Glückauf* 66, 1497-1502, 1538-41(1930)) has been improved to include hydraulic pressure measurement and other simplifications in recording, etc. The importance of coal moisture content, particle size and uniformity in filling the apparatus is stressed. Data are given to show the duplicability of results in tests. F. W. JUNG

The washing of coal on the Hoyer washer. ARTHUR GROUNDS. *Fuel Economist* 5, 561-4, 685-9(1930)—The Hoyer washer is based on a combination of the trough washer and the upward-current classifier. LESLIE B. BRACE

Incombustibles in coal and slag formation. REINHARD SCHULZE. *Die Wärme* 54, 81-5(1931)—Results of expts. are presented which indicate that in spite of practically the same ash content and heating value of the fuel, the combustion characteristics and especially the slag formation in large measure depend on the mineral chem. compn. of the incombustible matter and on the particle size, d and reaction velocity as well as on the kind of atm. in the burning charge. Greater Fe_2O_3 and CaO contents lower the slag m. p. of the fuel with the simultaneous presence of a mixed atm. and lead to the formation of a vitreous pasty slag. Coke in coarse granules can be burned without difficulty with a high output, while in smaller granules the same coke is almost unserviceable. ALLEN S. SMITH

Calorimetric investigations of slag. A. I. KORELIN. *Izvestiya Teplolekhn. Inst* (Trans Thermo Tech Inst Russia) 1930, No. 1, 22-8.—Calorimetric investigations of slag can be carried out in a bomb, a minimum of 15% of combustible material is required for an accurate detn. The combustion should be effected with an admixt. of an air-dry material low in ash and sulfur. Such material may be hard coal, anthracite, coke or wood charcoal. The bomb should be charged up to 40 atm. with O_2 and the slag should be moistened with mineral oil, crude oil, fuel oil, etc. Combustion in the presence of a mixt. of benzoic and salicylic acids or sugar generally does not give satisfactory results. The presence of some inorg. compds., such as a high percentage of carbonates, may affect the combustion values considerably. When burning a mixt. of slag and coal in the calorimeter, an excessive rise in temp. should be avoided. This would require a large amount of charging material, while working with small amounts produces sufficiently accurate results. A. A. BOGHTLINGK

Petrographic investigation of bituminous coal briquets. H. BODE. *Brennstoff-Chem.* 11, 476-8(1930), 12, 7-9(1931)—The briquetting of fine coals is discussed from the viewpoint of constitution, i. e., the durain, vitrain and fusain contents. Briquets were made (a) from equal mixts. of high volatile and low volatile coals of various particle sizes, and (b) from a mixt. of the high-volatile and a pressed, low volatile coal, the same amount of pitch was used in each case. Abrasion tests, as detd. in a tumbler, gave 22.5% of < 3 mm. size for the first mixt. and 15.5% of > 5 mm. size for the second. The effect of fusain content upon the pitch required is discussed at length, a contrast being noted with its effect upon the briquetting of brown coal. Distinction in fusain

effect is made for particle size, < 0.12 mm requiring an unnecessarily greater amount of binder pitch. Air cleaning does not satisfactorily reduce fusain content by removing undesired dust. For the various coal particle sizes of > 0.5 mm to 0.088 mm for the three coals used, analytical screen data are tabulated for fusain, vitrain, durain and mineral matter to demonstrate the increasing contents of the first two with increasing fineness. Magnified sections show the distribution of the aggregate in the briquets made.

Studies in the development of Dakota lignite. IV. Critical oxidation temperature of lignite. W. C. FAYON, G. A. BRADY, A. W. GAUGER, IRVIN LAVINE AND C. A. MANN. *Ind Eng Chem* 23, 87-93(1931).—The authors describe a method used in the detn of the CO_2 index and "critical oxidation temp" (C O T) of lignite. The results of the expts indicate (1) in a comparison of lignite with other coals, an increase in indexes and C O T with the rank of the coal, (2) a decrease in index and C O T with a decrease in the water content of lignite, (3) a decrease in C O T and an increase in index with decrease in particle size, (4) an increase in C O T and index as gases are changed from O_2 to air, to N_2 and to no gas (5) a decrease in both index and C O T with increase in gas flow (6) 80° to 90° as the temp at which, under the conditions used, dry lignite begins to self heat sufficiently to burn spontaneously, (7) no greater change in index and C O T after processing by the Heissner method than would be expected.

Carbonizing and briquetting Saskatchewan lignite. W. G. HERTINSTEAL. *Trans Can Inst Mining Met* 32, 39-40(1929).—A detailed description of the use of the Lurgi process for lignite contg 30% of moisture.

Fusain. S. W. PARR, H. C. HOPKINS AND D. R. MITCHELL. *Ind Eng Chem Anal. Ed* 3, 64-5(1931).—The modes of occurrence of fusain in Illinois coals are described. In selected lumps of freshly mined Ill coals 2 kinds of fusain, "hard" and "soft," were noted. Proximate analyses on moisture-free basis of 11 samples of fusain from 3 different beds of Ill coal and of face samples of coal from the same mines, and calcd values for "unit volatile matter" of the fusains are tabulated. The most striking feature noted was the fairly consistent character, especially in B. t. u. values, of the fusain from all parts of the state. The high porosity of fusain was shown by its air-drying to 0.5-2% moisture, while coals from the same beds retained 2-6% moisture. This porosity of fusain accounts for the large absorption of water by some bituminous coals even though they do not slack. The av "unit volatile matter" for the fusain was 20.44, which would appear to refute completely the forest fire theory of its formation.

X-ray study of vitrains. C. MAHADEVAN. *Indian J Physics* 5, 525-41(1930), *Fuel in Science & Practice* 9, 574-80, *et. C. A.* 24, 5952.—Paleozoic and Tertiary vitrains of varying compn were examd. by x-ray diffraction methods. The halo spacings for the former were 3.37 A. U. for the inner and 2.12 A. U. for the outer halo. For the latter, the values 3.50 A. U. and 2.21 A. U. were obtained. The intensity of general scattering between the direct spot and the halos depends somewhat on the moisture content and volatile matter of the vitrains. A calcn of the size of the diffracting particles indicates that they are of colloidal dimensions and that the mineral matter present is in a colloidal state.

Natural gas in 1929. G. R. HOPKINS AND H. BACKUS. *Bur of Mines, Mineral Resources of the U. S. 1929*, Pt II, 319-40 (preprint No. 22, published February 3, 1931).

Progress in the gas industry in 1930. HAROLD E. COFF. *Gas Eng* 48, 79, 80 (1931).

The manufactured-gas industry: its trend and problems. R. B. LECKIE. *Gas Age-Record* 66, 546-51(1930).—A review.

The trends of the natural gas industry. WRIGHT L. FELT. *Gas Age-Record* 66, 549-51, 558, 562, 595(1930).—A review.

Proposed table of standards for technical fuel gases. J. KRZYKIEWICZ. *Przemysl Chem* 14, 345-53(1930).—The 2 sheets of this table are reproduced and discussed. The first sheet shows the principle on which the gases are classified with respect to their origin, group, kinds and varieties, methods of prepn, brief descriptions of the varieties of the gases, and limits of their heat values. The second sheet covers addnl technical names of industrial gases.

The utilization of town gas as a fuel in heat-treatment furnaces. C. M. WALTER. *J Inst Fuel* 3, 408-15(1930).

Instruments aid in operating gas producers economically. JAY S. MCCLIMOV. *Steel* 87, 15, 54, 56, 58(1930).

F. W. JUNG

M. C. ROGERS

B. C. A.

W. W. HODGE

MALCOLM DOLE

E. H.

E. H.

R. B. LECKIE, Gas

LESLIE B. BRAGG

Gas Age-Record 66,

LESLIE B. BRAGG

J. KRZYKIEWICZ, Przemysl

Chem 14, 345-53(1930).

The 2 sheets of this table

are reproduced and discussed.

The first sheet shows the principle

on which the gases are classified

with respect to their origin, group,

kinds and varieties, methods of prepn,

brief descriptions of the varieties

of the gases, and limits of their heat

values. The second sheet covers addnl

technical names of industrial gases.

A. C. ZACHLIN

C. M. WALTER

J. Inst Fuel 3, 408-15(1930).

Instruments aid in operating gas

producers economically. JAY S.

MCCLIMOV

Steel 87, 15, 54, 56, 58(1930).

LESLIE B. BRAGG

Up-to-date methods for purifying illuminating gas before delivery to the consumers. A. V. SIVOLOBOV. *Izvestiya Teploelektr. Inst. (Trans Thermo-Tech Inst Russia) (Repts of the 5th All Russian Thermo-Tech Congress) 1930, No. 3, 73-100*—A general description of known methods. A. A. DOMITLIN

The Feld processes for the extraction of ammonia and hydrogen sulfide from coal gas. A. PARKER. *Gas Eng. 47, 694-5 (1930)*—P. discusses briefly Feld's processes for removing H_2S and NH_3 from city gas. The first process, consisting in the oxidation of H_2S by SO_2 , was not successful because of the slow rate of the reaction. In the second process $ZnSO_4$ (from ZnO and SO_2) was used to oxidize the H_2S , the products (ZnS , S and water) could be treated with SO_2 to form ZnS_2O_3 , and the mixt. again used to treat the gas. This process was unsatisfactory because of side reactions. The third process involved the removal of both H_2S and NH_3 by reaction with $FeSO_4$ to form $(NH_4)_2SO_4$, and further treatment producing Fe polythionates which decompose in the presence of water and NH_3 . The difficulties with the above processes resulted in the development of a process using NH_3 polythionates for removing H_2S and NH_3 simultaneously. Also in *Inst. Gas Eng. Communication No. 10, 4-8 (1930)*.

M. C. ROGERS

Determination of carbon monoxide in illuminating gas. JOSEF DUBOIS. *Przemysl Chem. 14, 313-8 (1930)*—From a crit. exam. of the common methods of analysis D. finds those using Cu salts to be the best suited. The acid soln. of Cu_2Cl_2 absorbs CO rapidly but incompletely, while the ammoniacal soln. absorbs it completely though slowly. The best results are obtained by combining the 2 kinds of absorptions, allowing the acid soln. to take up 80% of the total CO content and removing the remainder in the ammoniacal soln. The method consists in first shaking the sample 2 min. with each of 2 successive 10 cc. quantities of the acid reagent, and then for 2 min. with each of 3 successive 5 cc. portions of the ammoniacal reagent. A complete absorption of CO takes place. The Damiens reagent was adopted for a complete and rapid absorption of CO over H_2 . Three cc. of the reagent is introduced into the H_2 pipet and shaken lightly for 9-10 min. Methods of prep. the solns. are given. A. C. Z.

Atomizing and wick odorizers for natural gas. PHILIP E. BECKMAN. *Gas Age Record 66, 850-64, 886 (1930)*—A Venturi type atomizer and a wick odorizer for odorizing natural gas are described in detail. Considerable exptl. work led to the choice of the present types of app. LESLIE B. BRAGO

New aspects of the gum problem. W. L. SUTVELY. *Gas Age Record 66, 577-8, 584 (1930)*—Gum troubles have recently been noted with coal gas in pilot needle valves and thermostats. The gum is carried to the point of deposition as suspended particles of very small size. Preventive measures have been successfully applied but differ for individual situations. LESLIE B. BRAGO

Gum formation in town gas. L. H. SENSIBLE. *Gas Eng. 47, 625-7 (1930)*—Gum forms in small droplets, continuous impacts against pipe walls and fittings assist these to deposit and form large drops. Volatilization of the lighter components causes the gum to become less fluid with increase in time. Gum deposits on valves in meters and deteriorates meter leathers. It stops up orifices in burners. Causes of gum formation and requirements of effective inhibitors are discussed. M. C. ROGERS

Flue-gas testing. H. D. BRASCH. *Die Wärme 54, 57-9 (1931)*—A mech. Orsat type of continuous flue gas analyzer is described. The percentages of CO_2 and combustible gases are indicated and are recorded on a strip chart. Details of the construction of the instrument and of its installation are given. ALLEN S. SMITH

Combustion temperature and flue-gas composition; their determination with regard to dissociation. ROMAN SELIKIN. *Die Wärme 54, 61-3 (1931)*—A method of detg. the flue-gas compn. and combustion temp. of any gas mixt. by means of the chem. equil. const. is discussed. When the gas contains no CH_4 , the calcs. are simple. If CH_4 is present, the method is applicable if the fact is taken into consideration that CH_4 is not stable at high temps., but dissociates into its constituents which burn to CO_2 and H_2O . The combustion of an illuminating gas with $2/3$ the theoretically required air is used as an example. ALLEN S. SMITH

Determination of viscosity on small samples of tar. P. M. POTTER AND J. S. SACH. *Roads 8, 231-2 (1930)*—The app. described by Mallison (*C. A. 15, 1231*) is shown to provide a rapid means of checking the viscosities of tars, and the precautions necessary to obtain consistent results are discussed. B. C. A.

The thermal decomposition of the low-temperature tar of Fushun coal. I. TADA-SHI. *Mizoshita Abstracts from Repts. Central Lab. S. Manchuria Ry. Co. 1929, 33-5*—A yield of 13.5% anti-knock gasoline is obtained by cracking at 430° and 25-30 atm. V. F. HARRINGTON

The utilization of high-temperature coal-tar pitch. I. Preparation of creosote-oil substitute from pitch. MUNENARI TANAKA, KILO-SHI MORIKAWA AND ISAO MORIKAWA. *Abstracts from Rept. Central Lab. S. Manchuria Ry. Co.* 1929, 29-32.—Gasoline solvent naphtha mixts (1:1) dissolve 46% of the pitch. This ext. mixed with creosote oil is used as a creosote-oil substitute. II. The preparation of active carbon from the extraction residue of high-temperature pitch. SUSUMU WATANABE. *Ibid.* 32-3. Calcining with K_2SO_4 gives a C more active than any commercial product. V. F. HARRINGTON.

The desulfurization of brown coal tar oils with brown coal low-temperature coke. C. STAENGLER. *Brennstoff Chem.* 12, 43-5 (1931).—Methods proposed and developed during the past year are reviewed. Attempts were made to desulfurize a generator light oil of 0.46% S, its distillate of 0.42% S, and a 'yellow' oil of 0.72% S by distn. over dried coke of 3-4 mm size in an electrically heated tube furnace. While the H₂S from S compds. decomposed at the higher temps. was removed completely, the S removed was consistently small with a max. of less than 40% of the total S. Distn. in a stream of air or H had no effect. Pretreatment of the coke with HCl did not effect removal of the S, but it discolored the distillate with colloidal matter. Pptn. of $CaCO_3$ in the coke increased the S removal to 50% but is not believed practical. F. W. J.

The agglutination of coal and the activation of its surface during coke formation, considered as two complementary phenomena. W. SWIETO-ŁAWSKI. *Fuel in Science & Practice* 9, 564-6 (1930).—The necessity of thorough impregnation of an inert material with caking material if the agglutinating ingredients of the latter are to be used effectively is stressed. The formation of structured surfaces on the stable products of dry distn. is held to be a converse process to that of the production of well-cemented coke.

D. A. REYNOLDS

A comparison of methods for testing the caking properties of coal. L. SLATER. *Fuel in Science & Practice* 9, 580-91 (1930).—See C. A. 24, 3325. D. A. REYNOLDS.

Coking a banded bituminous coal: the part played by each band. C. P. FINN. *Trans. Inst. Mining Eng.* (London) 80, 283-302 (1931).—Exptl. results show that different samples of clarain, which forms the largest percentage of the banded ingredients present in coking slack, may vary considerably in their coking properties but yield a comparatively poor coke that is much fissured and breaks easily. Evidence available indicates that vitrain is proportionately more valuable than clarain in a coking slack. Durain appears to exert little effect. Fusain, although itself non-caking, has been shown to be valuable when uniformly admixed with the remainder of the coking slack in amts. < 5%. An increase in the amt. of fusain middlings does not have any beneficial effect, and there is an indication from the tests that their complete removal may be advantageous. The presence of free dirt is deleterious to the quality of the coke obtained.

C. W. OWINGS

Production of low-sulfur coke in the Ekaterinov ore area. V. M. STAKHOVSKII AND R. M. IL'YA RAITNER. *J. Chem. Ind. (Moscow)* 6, 734-5 (1929).—Crushed coal from the Almaz vein graded on a 3 mm screen was found to show lower ash and S contents in the fines than in the coarse material.

A. C. ZACHLYN

The dry cooling of coke. JENS RUDE. *Engineering* 130, 543-4 (1930).—The thermal advantages of dry-cooling coke are outlined, and it is stated that the recoverable sensible heat is slightly less than 4% of the B. t. u. of the coke taken as 11,000 B. t. u. per lb. Section diagrams of the Salter dry coke-cooling plant are given. In this process inert combustion gases circulate through the hot coke and to a steam generator all in a closed system. Cooling time varies from 3.5-4 hrs. for a large capacity plant to 1.5 hrs. for small gas works where the car of coke (1.5 tons) is wheeled directly into the cooling chamber. The essential features of the Collin dry-cooling plant are also outlined. The advantages of dry coke-cooling are: better quality coke, no water to evap. when the coke is burned, a smaller amt. of fines and breeze and recovery of the sensible heat of the coke. The chief objection is the high cost for a plant of the large dimensions required because of the comparatively low temp. gradients existing throughout the process. A set of curves is given showing the relative heat-transfer rates from coke to boiler water. Apparently a saving of approx. 1/3 in the cost of the plant could be secured by cooling the coke to only 450-500°, but some 20% less steam would be produced than by cooling to about 300° as is now done. Two systems using steam (C. A. 24, 5466) in place of the inert gases are outlined and their relative advantages discussed.

W. W. HODGE

Determination of volatile matter in coke. A. P. SHAKHOV AND M. D. ZHUKOV-SKAYA. *Izvestiya Teploelektr. Inst. (Trans. Thermo-Elec. Inst. Russia)* 1930, No. 2, 17-26; cf. C. A. 24, 1721.—Results obtained by the use of alc., gas and elec. heating equipment are compared.

A. A. BOHRLING

The use of anhydrite (calcium sulfate) in the production of ammonium sulfate. P. PARRISH *Gas Eng* 47, 693-6 (1930), cf C A. 22, 4768 —P discusses the production of $(\text{NH}_4)_2\text{SO}_4$ by the action of $(\text{NH}_4)_2\text{CO}_3$ on CaSO_4 . The carbonation and reaction with CaSO_4 must be carried out in sep. vessels. This increases the reaction rate and prevents side reactions. Four vessels in series are required. Also in *Inst Gas Eng., Communication No. 19, 9-24 (1930)*. M. C. ROGERS

Rapid method for determination of benzene and phenol in ammoniacal and waste liquors. W. MUNZ *Brennstoff-Chem* 12, 3-4 (1931). —To det. C_6H_6 , pass upward through a vertical tube, 30 cm long and 5 cm in diam, contg. 100-150 g activated charcoal, 10 l of the liquor to be tested in 1 hour's time. Drain the charcoal, recover the C_6H_6 by means of superheated steam by the method of Kattwinkel (cf C A. 24, 1252) and det. to 200° by the usual procedure. This method is accurate only up to a charcoal absorption of 15% C_6H_6 by wt. To det. PhOH measure into a 250-cc distn. flask 25 cc of the filtered liquor. Add ammoniacal CuSO_4 soln (prepd by adding concd. Ni_2OH in excess to cold satd. CuSO_4 soln) in excess of pptn. and then 5 cc 40% NaOH soln. Boil over a free flame with shaking until NH_3 vapors are no longer evident, connect the flask to a condenser and pass CO_2 through the soln. at the rate of 2-3 bubbles per sec. The neutralization should be complete after some 5 min. Distil to dryness with a slow stream of CO_2 . Rinse the condenser, adapter, etc., into the 250-cc volumetric flask used as receiver and make up the distillate to vol. For titration, 25 or 50 cc of this is used in the case of NH_3 liquors, but with waste liquors the whole may be necessary. To this in a stoppered Erlenmeyer flask add 25 cc. of 0.1 N $\text{KBr}-\text{KBrO}_3$ soln and then 15 cc 15% H_2SO_4 . After $\frac{1}{2}$ hr add 10 cc of 10% KI soln, let stand $\frac{1}{4}$ hr, and titrate the excess using starch indicator, with 0.1 N $\text{Na}_2\text{S}_2\text{O}_3$ soln. F. W. JUNG

The selection of coke for foundry purposes (HUPFER) 9. Ashes and the remains of fuel found by the excavation of ruins of a seventeenth century glass factory near the Keizergracht at Amsterdam (KREULEV) 2. Colorimetric determination of the sulfate ion in coal (GUARNIER) 7. Geological aspects of the formation of coal (FOX) 8. Development of Dakota lignite (HARRIS, et al) 8. Alkali ferrocyanides [preparation during extraction of HCN from coal gas by washing] (Fr. pat. 692,868) 18. Separating gases by liquefaction (Brit. pat. 335,509) 13.

RIGGA, FLORIAN. Die Kohle, ihre Entstehung und ihre Verwertung. Prachatische Schramm 15 pp. K6 3

Fuel briquets. H. J. M. WASTEELS. Belg. 371,152, July 31, 1930. The coal is placed in iron or steel molds which are then placed in the distg. retorts. After distn. the balls, briquets or blocks of coke are demolded.

Fuel briquets. LUDWIG WEBER. U. S. 1,791,077, Feb. 3. Particles to be briquetted such as fuels, ores or fine dust are mixed with very finely divided clay of great plasticity and this mixt., before briquetting, is moistened with sulfite waste liquor.

Production of agglomerates without the addition of pitch. H. HARDY. Belg. 370,410, June 30, 1930. Coal or lignite is prepd. in a globular state with a view to soln. in solvents acting successively, e. g., the globular coal is treated with heavy oils, solvent naphtha and benzene, either in the above or reverse order, the coal can be carbonized at low temp., to obtain a powdered fuel, or it can be first pressed to agglomerate it and then carbonized for the production of semi-coke. Cf C A. 24, 4615.

Plant for the manufacture of smokeless fuel. COMPAGNIE DES MINES DE VICOIGNE NOEUX ET DROCOURT. Fr. 693,273, May 30, 1929.

Fuel for internal-combustion engines. JOHANN W. MICHEL. Ger. 515,076, Jan. 9, 1929. Soot from solid, liquid or gaseous fuels is compressed into briquets and then finely ground.

Motor fuel. I. G. FARBENIND. A-G. Fr. 692,594, Mar. 24, 1930. When starting liquid fuel of high vapor tension is led to explosion motors by mech. pressure on spongy material contg. the fuel.

Gasifying wet bituminous fuels. HERMANN HILDEBRAND. U. S. 1,791,411, Feb. 3. Alternately operated regenerative chambers are heated by burning a portion of recirculated gas, a mixt. of recirculated gas and steam is superheated in one of the heated regenerative chambers and thence passed upwardly into a gasification zone of a fuel bed to form water gas, a portion of the water gas and recirculated gas is passed through a superposed distn. zone of the fuel bed, effecting distn. of gases and vapors from the fuel, another portion of the water gas and recirculated gas is passed in heat exchange

with the wet fuel and to a collecting main, and the distn gases and vapors are recirculated through the regenerative chamber. App is described.

Low-temperature carbonization of bituminous carbonaceous materials. I. G. FARBENIND. A.-G. Brit. 335,740, Sept. 30, 1929. In a process such as described in Brit. 301,975 (C. A. 23, 4323) in which hot gases are blown through fuel so that it is maintained in active motion on the grate, to effect carbonization without combustion or gasification, impact devices are arranged over the fuel bed to limit the upward movement of the fuel. Various details of app are described. Cf. C. A. 24, 459.

Low-temperature carbonization of powdered fuel. F. L. DUFFIELD. Brit. 335,643, July 5, 1929. In the use of a horizontal retort using steam as the carrying medium, the retort is locally heated to a high temp, and steam is admitted tangentially to the heated surface. Various details of structure of the app and of its operation are described.

Process and oven for the distillation or low-temperature carbonization of coal and other carbonaceous materials. C. HONNAV. Belg. 370,260, June 30, 1930. Constructional features.

Vertical retort for low-temperature carbonization of fuel. CHARLES TURNER. Ger. 515,388, Mar. 20, 1928. See Brit. 295,461 (C. A. 23, 2276).

Distilling solid fuels. FRANK E. HOBSON. Ger. 515,386 and 515,387, Feb. 3, 1927. See Brit. 277,214 and 277,215 (C. A. 22, 2155).

Distilling solid carbonaceous materials in oil. IMPERIAL CHEMICAL INDUSTRIES, LTD. Ger. 515,389, Oct. 28, 1928. See Brit. 305,741 (C. A. 23, 5032-3).

Furnace for the low-temperature distillation of carbonaceous materials. CHARLES HONNAV. Fr. 693,038, Mar. 29, 1930.

Rotatable retorts for the low-temperature distillation of coal. THE CARDOCITE CO. Fr. 692,881, Mar. 28, 1930.

Apparatus for distilling coal for the production of coke. BABCOCK & WILCOX, LTD. Fr. 692,520, Mar. 21, 1930.

Apparatus for carbonizing comminuted coal and for enriching water gas with distillation products. EDWARD A. DIETRICH. U. S. 1,700,745, Feb. 3. Structural features.

Destructive hydrogenation of coals, tars, etc. N.-V. DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ. Brit. 335,543, June 21, 1929. Various carbonaceous materials, which may be mixed with phenolic substances or like O contg. compds are first treated with CO or gases contg. CO at elevated temps and preferably under pressure to effect removal of most of the O present, and are then subjected to a treatment with H₂. Water present in the initial material may react with the CO to form H₂ and effect partial hydrogenation. Various details of procedure, catalysts used, etc., are given. Cf. C. A. 25, 1362.

Purifying gas. HUMPHREYS & GLASGOW, LTD. Fr. 693,100, April 1, 1930. Gases are freed from S by bringing the gases into contact with a heated catalyst formed of ZnO with or without an activating agent, whereby the org. S compds are decomposed forming H₂S which is eliminated. The undecomposed org. S compds are removed by H₂SO₄. Cf. C. A. 24, 1726.

Desulfurizing gases. RICHARD BRANDT. Ger. 514,666, Dec. 15, 1927. Addn to 503,118. In removing H₂S from gases by means of K₂I c(CN)₂ soln which is regenerated as described in Ger. 503,118 (C. A. 24, 4991), the desulfurizing soln is used in a continuous cycle without filtering off the free S liberated from the gases, the cycle including a settling vessel in which the agglomerated portion of the originally colloidal S is sep'd.

Dehydrating fuel gas. FREDERICK W. SPERR, JR. (to Koppers Co.) U. S. 1,701,080, Feb. 3. Gas such as coke-oven or water gas is compressed, cooled by direct contact with water at substantially atm. temp. so that a portion of the moisture carried by the gas is removed, and the compressed and cooled gas is then treated with a hygroscopic material such as CaCl₂ or H₂SO₄ for further removal of moisture. An arrangement of app is described.

Gases containing hydrogen. I. G. FARBENIND. A.-G. Brit. 335,524, April 24, 1929. In the conversion of carbonaceous materials such as C₁₁, coal distn. gases, tars, mineral oils, etc., into gases consisting mainly of H₂ and CO or of these gases and N₂, by treatment with air, O or CO₂ (with or without steam) in an elec. arc furnace, devices such as narrow exit openings produce eddies which cause all parts of the gas mixt. to attain a temp. of at least 1400°. Various details of app and procedure are described.

Oil gas. CONSTANTIN CHILOWSKY. Ger. 514,487, Mar 25, 1928 See Fr. 679,107 (C. A. 24, 3888)

Mixed oil gas and water gas. H. G. TERZIAN (to Humphreys & Glasgow, Ltd.). Brit. 335,495, July 11, 1929 In a process in which oil gas is passed through an ignited fuel bed and the deposited C is consumed in the water gas reaction, steam is generated in the tar washer by the passage of hot oil gas, and the steam thus formed is used in the generator to form water gas. Various details of app and procedure are described. Cf. C. A. 25, 1063

Water gas. HUMPHREYS & GLASGOW, LTD. Fr. 603,105, April 1, 1930 Water gas is made by passing steam upwardly through a mass of incandescent fuel from a point situated sufficiently above the base of the mass to leave an annular zone of incandescent fuel. A part of the gas produced is raised through a superposed layer of fresh fuel to eliminate hydrocarbons, another part is passed into a carburetor where it is mixed with a hydrocarbon, and the mixt is introduced at the center of the lower part of fresh fuel.

Water gas. POWER GAS CORPORATION, LTD., and NILES E. RAMBUS. Fr. 603,514, Apr 8, 1930 Means is described for superheating the steam, which is passed alternatively in an ascending and in a descending direction through the fuel. Cf. C. A. 24, 3888

Water gas. HENRY O. LOEBELL (to Henry L. Doherty). U. S. 1,700,824, Feb 3, 1928 Fuel is passed downwardly in an unobstructed column of circular cross section through a gas generating zone in which high temps are maintained by periodically blasting air in a substantially radial direction through incandescent fuel in the zone. Steam is passed through incandescent fuel in the high temp zone during periods following the air blasting, and hot water gas thus formed is passed through fuel above the blast zone. App is described. Cf. C. A. 25, 582

Apparatus for manufacture of water gas. HENRY O. LOEBELL (to Henry L. Doherty). U. S. 1,700,823, Feb 3, 1928 Structural features.

Water-gas producer. SOC. ANON. LA CARBONITE. Fr. 603,280, July 8, 1929 Device permitting of the use of high-volatile or mixed high- and low-volatile coal fines in gas producers. E. DOTTRE. Belg. 370,238, June 30, 1930 A circular rotating plate is placed above the producer. The fuel is fed on to this plate and allowed to stay there sufficiently long to become agglomerated before it is fed into the producer. The gases given off from the coal while on the plate are mixed with the producer gas.

Apparatus for producing coal gas and water gas from coal dust. KARL MEITZLER. Ger. 514,486, Sept 4, 1925

Ovens for producing gas and coke. STETTINER CHAMOTTE FABRIK A. G. NORM. DIN. Ger. 515,385, Dec 11, 1928 Constructional features are described.

"Waterless" gas holder. R. & J. DEURSTEN, LTD., J. W. SCOTT and WATERLESS GAS HOLDER CO., LTD. Brit. 335,502, Feb 15, 1930

Purification and desulfurization of benzene. C. PETIT. Belg. 371,908, Aug 31, 1930 The benzene vapors are bubbled through liquors such as H_2SO_4 , NaOH or Na_2CO_3 , which may be made to flow over the trays of a benzene rectifying column or of a column placed after the rectifying column.

Storing acetylene. F. ROFFEY and IMPERIAL CHEMICAL INDUSTRIES, LTD. Brit. 335,820, Nov 28, 1929 C_2H_2 is stored in cylinders or other containers contg a solid absorbent material impregnated with an C_2H_2 -solvent comprising a relatively non-volatile ether, ester or ether-ester of a polyhydric alc. such as glycol diethyl ether, glycerol triethyl ether, glycerol diacetate, glycerol triacetate, β -ethoxyethyl acetate, β - γ -diethoxy n -propyl acetate, glycerol monoacetate, glycerol diacetate, glycerol monoethyl ether and glycerol diethyl ether or mixts of these. Before storing, the C_2H_2 is preferably dried as by passing over $CaCl_2$ before and after compression. Some other suitable solvents are also mentioned.

Phenols. HUILES, COUDRONS ET DÉRIVÉS and GEORGES SUPAIN. Fr. 603,734, July 22, 1929 Phenols obtained in phenolic oils and tars are extd by solns of alkali carbonates at a high temp under pressure, means being provided to remove the CO_2 formed. This CO_2 is used for acidification of the phenolates obtained to reform carbonates.

Tar refining. JEAN H. BREGI (to Compagnie Technique des Pétroles). U. S. 1,701,052, Feb 3, 1928 See Fr. 680,204 (C. A. 25, 1064).

Apparatus for distilling tar, oils, etc. COMPAGNIES RÉUNIES DE GAZ ET D'ÉLECTRICITÉ. Fr. 693,681, April 10, 1930

Coking pitch. C. OTTO & CO. G. M. & H. Ger. 514,478, Feb 10, 1928 A finely

ground mixt. of pitch and pitch coke is coked in an oven of refractory material, the walls of which are coated with pitch coke to prevent penetration by the pitch.

Coke oven and door. CARL STILL. U. S. 1,700,775, Feb. 3. Structural features.
Method and plant for dry-cooling coke. WILHELM KLEISSER. Ger. 514,410, Nov. 18, 1925.

22—PETROLEUM, LUBRICANTS, ASPHALT AND WOOD PRODUCTS

W. F. PARAGHIAN

Municipal regulation of oil storage. ROBERT S. MOULTON. *Quart. Nat. Fire Protect. Assoc.* 24, 255-76 (1931). A thorough review of the fire hazards of petroleum products and their control in cities, accompanied by a list of N. F. P. A. publications bearing on the subject. C. I. JONES.

Catalytic reactions of sulfur compounds present in petroleum. I. High-sulfur naphthas in contact with nickel and iron catalysts. J. C. ELLAN, G. H. WILDER AND H. S. TAYLOR. *Ind. Eng. Chem.* 22, 1281-90 (1930). Reduced Ni removed S from naphthas in both the liquid and the vapor states. The catalyst was most effective in its action upon the first of the naphtha vapor passed over it. Its activity then dropped off till a "steady state" was reached and its effect was almost constant for the remainder of the run. The S content of the naphthas was reduced still more if H_2 was passed with the naphtha vapor over the Ni. The action of reduced Fe was similar to that of Ni but less pronounced. II. Pure sulfur compounds in hydrocarbon materials in contact with nickel catalysts. J. C. ELLAN. *Ibid.* 1290-3.—Sols. of $BuSH$, $iso-BuSH$, Pr_2S , $(iso-Bu)_2S$ and thiophene in naphthas were passed in vapor form over reduced Ni catalysts at 300°. The mercaptans readily decomposed giving up S as H_2S . The sulfides reacted less readily. Thiophene was not affected after the catalyst had reached the steady state unless H_2 was passed with the naphtha vapor. Thiophene is formed in the thermal decomposition of mercaptans and sulfides, hence the removal of S from cracked naphthas is likely to be more difficult than from straight run. FAYE B. CRANDALL.

Refining of light petroleum distillates. H. P. RUP AND RALPH H. PISACH. *Pur. Mines, Bull.* 333, 111 pp (1930).—Proper fractionation of the gasoline from a pressure distillate will decrease the amt. of chem. treatment necessary to make the gasoline meet trade requirements. To get the max. benefits of fractionation on the chem. treatment, a good degree of sepn. is necessary; there should be no overlaps in distn. range. In fractionation units, a reflux (closed-coil or open-spray) at the top of the bubble tower is necessary to control the top temp. When a closed-coil reflux is used, incoming charging stock may be preheated by pumping it through the coil. When an open spray is used, the refluxing medium should be the material that is taken overhead as a vapor. The present methods of treating light distillates are described in detail. The influence of time, temp. and thoroughness of contact in treating with H_2SO_4 is considered. The effects of acid treatment on C deposits, distn. range, gravity and color are shown. The use of alkali is discussed. Studies with S and S compounds dissolved in gasoline are described. The effect of fractionation on color and gum formation is discussed. ALDEN H. FERRY.

Refining cracked distillates. JACQUE C. MORRELL. *Natl. Petroleum News* 22, No. 51, 61-6, No. 52, 49-51, No. 53, 46-8 (1930); 23, No. 1, 63-72, No. 2, 67-8, No. 3, 51, 54, 58, 60 (1931).—The following are discussed: the desulfurization of cracked distillates by different methods; theory of gum formation, use of gum inhibitors for gasoline to be stored; H_2SO_4 treatment, split plumbite treatment; refining by sludge acent, recovery of spent chemicals at the refinery; acid and "doctor" soln.; rerunning acid treated distillates, soap washes, vapor-phase treating, treatment of end point cracked distillates without rerunning; continuous treating plants, mixing devices, settling tanks and sepg. devices; gasoline treating plants; distn. equipment, operation of shell stills; operation of shell still rerun units, pipe still rerunning, operating precautions, chem. factors in distn.; future possibilities. R. W. KELLY.

The occurrence of higher fatty acids in mineral oil distillates. E. HOLTSMANN AND STANISLAUS VON PILAT. *Brennstoff-Chem.* 12, 41-2 (1931).—The cryst. material previously (cf. C. A. 25, 581) sepd. from the phenol free naphthenic acid fraction of petroleum has been identified as higher fatty acids, although work by others indicated their absence in Polish oils. More of this material was prepd. from a Boryshiv machine oil by dehydrating and distg. the acidified Na_2CO_3 washings to 275-40° at 12-15 mm. Hg in a

new Burstin Winkler app., shown schematically, and sepd the distillate into the acids, phenols and hydrocarbons. The first fraction yielded the fatty acids upon cooling. After removal by centrifuging, recrystn from MeOH and then from acetone, white crystals were obtained. A mixed m p identified them as arachic acid. This characterization was confirmed by the neutralization no., the mol wt. by m p depression and the elementary analysis. Additional fatty acids sepd from the mother liquor as Li salts give a total of 0.0013% for the crude petroleum. They are being investigated further. F. W. JUNG

Present-day tendencies in oil-fuel burning. JOHN L. STREYENS *Fuel Economist* 5, 569-72, 623-4 (1930).—A review. LESLIE B. BRAGO

The value of ash determinations of mineral and fatty oils. W. SCHAEFER *Seifen- und Ziegler-Ztg* 57, 878-9 (1930).—The org. S in oils usually enters the ash on burning, indicating the presence of sulfates or sulfides in the oil. An extn with H_2O or acid does not furnish evidence of the presence of these compds. P. ESCHER

Crystallization of paraffin wax. E. KATZ *J. Inst. Petroleum Tech* 16, 870-88 (1930).—The cryst. form of Polish and Asiatic wax crystd from various solvents is shown by a microscopic study to be fundamentally a six sided plate. Sixty two photographs illustrate the variations of this form. L. W. T. CUMMINGS

Evaluation of anti-knock properties of fuels. L. AUES *Automobiltech* 2, 33, 538-9 (1930).—A review. A. A. BOEITLINGER

Velocity of propagation of flame and pressure wave: thermodynamic consideration of motors and the phenomenon of detonation. E. VARETON AND U. RE *Ann. scuola ing. Padova* 4, 226-40 (1928).—A review of the literature and of the work done, chiefly by American and English workers. S. L. B. EDITION

Experiments to determine velocities of flame propagation in a side valve gasoline engine. H. S. GAYNE *J. Inst. Petroleum Tech* 16, 750-82 (1930).—It is known that small head clearance in an explosion engine tends to reduce detonation and that differences of head clearance cause variations of turbulence. The possible relation of velocity of flame travel to head clearance and to turbulence forms the subject of this study. Measurements of the velocity of flame movement were taken in an engine with side-valve low turbulence combustion chamber, and in 3 turbulent type Ricardo engines with head clearances of 0.022, 0.084 and 0.272 in. Castings of the several combustion chambers used could be bolted down to a single cylinder block. Six eight holes in a straight line were made in the top of each combustion chamber and fitted with glass windows. A disk 16 in. in diam revolved in a horizontal plane over the combustion chamber, on a vertical shaft connected with the ignition timing shaft. Its radius was such that the line of holes cut in it from center to circumference would coincide with the line of openings into the combustion chamber. Its timing was adjustable by a screw. With the engine running, the usual readings of speed, torque, fuel consumption, ignition advance and jacket temps. were made at the same time as the observations through the stroboscopic disk. By the use of the adjusting screw, the time at which the flame passed each window could be known in terms of crankshaft degrees before or after the top dead center. Pressure indicator diagrams were made. All runs were at 900 r p m. Max. velocity of the flame was reached in all cases within 7° and 10° after the top dead center. This was also the zone of max. rate of pressure rise. The smaller the head clearance, the more rapidly the velocity of the flame fell off after reaching it. This was attributed to the higher surface-vol. ratio by which the flame was cooled and its travel impeded. The same cause raised the H U C R also. The mean velocities in the 4 heads were proportional to the rate of pressure rise, which is taken as the measure of turbulence. The max. velocities were nearly proportional, with the exception of that in Head No. 2. The velocity with which the flame spread from the spark to the first window was also proportional to the rate of pressure rise or turbulence. EMMA E. CRANDAL

Synthetic lubricating oils. A. W. NASH, H. M. STANLEY AND A. R. BOWEN *J. Inst. Petroleum Tech* 16, 830-40 (1930).—Expts. are reported on the polymerization of C_2H_4 to light and heavy oils by heating in an autoclave at pressures up to 60 atm. Without a catalyst the polymerization occurred only above 325°. The temp. was lowered to 230-275° with $ZnCl_2$ as a catalyst and to 180° with $AlCl_3$, which was the principal catalyst studied. The product consisted of free oil (I) and hydrocarbon associated with the $AlCl_3$ (II). (I) contained from 10 to 45 C atoms and was of a satd character, while (II) contained 10 to 50 C atoms per mol. and was unsatd. The products were distilled and some of the properties of the fractions reported: b p., Br no., d., viscosity, and change of d. on oxidation. The heavier fractions had practically the

same characteristics as com lubricants, except that they were more readily oxidizable

L W T CUMMINGS

Recent examinations of friction and lubrication. J E Southcombe *Teknisk Ukeblad* 75, 351-8(1928) —A review with subsequent discussion C A RONAK

About germ process oil. SVFRER LAUNY *Teknisk Ukeblad* 75, 388-9(1928) —A reply to Southcombe (see preceding abstract) Since the fatty acids which are used to make germ process oil attack metals slightly with the formation of soaps, L asks for more detailed information, particularly in regard to the variation of the coeff of friction with time C A RONAK

The stability of germ-process oils. JAMES E Southcombe *Teknisk Ukeblad* 75, 407-8(1928) —A reply to Launy (cf preceding abstr) The fatty acids used for germ process oil in amts of about 0.7-1.0% are very pure, and accordingly have only a very small chem activity The standard test for germ oil is to heat 200 g of oil to 80° for 36 hrs in contact with steel plates No change in color or wt of the metal takes place Comparative tests with germ oil and mineral oil carried out by heating both to 80° with Cu foil for 36 hrs, show that the wt of the Cu foil is reduced a little more in germ oil than in pure mineral oil The reduction is still higher in compound marine oil contg rapeseed oil, which has been used extensively in ships for more than 2½ years C A RONAK

The stability of germ oils SVFRER LAUNY *Teknisk Ukeblad* 76, 10 1(1929) —A reply to Southcombe (cf preceding abstr) Even the tables given by Southcombe show that the action of "germ" to an oil always increases the corrosive action of the oil The amt. of Cu dissolved was increased by 58 and 50% in 2 expts C A R

Dielectric losses of oils. A GEMANT *Z. tech. Physik* 11, 514-5(1930) —By recent articles of G (Wiss Veröffentlich. Siemens Konzern, in press) and of Ornstein and Willemse (C A, 25, 9), it is shown that the loss curves for transformer oils consist of 2 superimposed parts (loss angle, δ , versus frequency) For high frequency the Debye losses are most important, but at low frequency the Wagner effect preponderates B J C VAN DER HORVEN

Determination of the drop point by means of the Ubbelohde apparatus, especially for high-temperature lubricants. P. SCHWARTZ *Svensk Papperstidn. 57*, 877(1930) —The detn of the drop point for greases and solid lubricants by means of the Ubbelohde app may give values that differ by 25°, due to differences in the diam or length of the glass tubing and to the depth to which the thermometer bulb is immersed The rate of heating and moisture in the high-temp lubricants also influences the results P FENNER

Improved asphalt extractor. T. E. STANTON. *Eng News-Record* 106, 38(1931) —A new type of extractor is described in which the solvent is passed through an aluminum filter stone 1 in thick with the aid of 90 lb air pressure The max variation between the asphalt content used and extd from 11 prep'd samples contg a high % of 200-mesh material and asphaltic cement was 0.2%, the av variation being less than 0.03% The solvent used is CHCl_3 , approx 85% of it being recovered The cost of solvent per sample is about 13¢ The time of extn with oil or asphaltic concrete mixts is 15-30 min R L THOMPSON

Filtration of asphaltic precipitates by suction, permitting rapid determinations. PAUL WOOD AND JEAN GIVAUDON *Bull soc chim* [4], 47, 1449-50(1930) —Powd glass that passes through a sieve of 0.054 mm mesh is used in a crucible with sintered-glass bottom for filtering asphalt The process is rapid and gives const results G M MURPHY

A study of gum turpentine. (MISS) MARCELLE BARRAUD *Bull inst. pin* 1930, 217-21 —The yield of spirit of turpentine from gum produced by pines of the same species or of species yielding spirit of turpentine having approx the same compn is fairly uniform, provided that the trees are situated in the most favorable conditions for growth Spirit of turpentine consisting of pinene and noplene can be obtained in yields exceeding 25%, but usually 22-24% Spirit of turpentine contg other terpenes than pinene and noplene is obtained in somewhat lower yields (20% or less). The first gatherings of gum contain more spirit of turpentine than the later ones Trees of large diam tend to give more gum contg a higher proportion of spirit of turpentine than small trees Analysis of 300 samples, each representing at least 1 tank-car of spirit of turpentine from Bordeaux pine, gave α_1 —28.98° to —32°; spirit of turpentine from Austrian black pine had α_1 —15.83° to —33.27°. α_2 was for Austrian pine and maritime pine, resp., —17.50° to —37.14°, +10.12° to —38.80° By fractionation of highly l rotatory maritime pine spirits of turpentine in a 2 m Lebel column, pinene was sep'd having $[\alpha]$ —51.13°, while the highest values previously reported by Dupont

were $+50.92^\circ$ for pinene from Aleppo pine and -48.70° for pinene from maritime pine. A similar study of spirits of turpentine from Aleppo pine would probably yield a pinene having $[\alpha]$ higher than $+50.92^\circ$.
A. PAPIEBAU COUTURE

Two important physical properties of fuels, pertaining to the formation of mixtures in carburetor engines (WAWAZINTOX) 21. Petroleum insecticides (WOODWORTH) 15. Limestone oil reservoirs of the northeastern United States and of Ontario, Canada (MURRAY) 8. Some properties of limestone as a reservoir rock (HOWARD, LOVE) 8. Study of the black shale overlying the cap rock of the Cromwell sand in relation to the origin of the Cromwell oil dome, Oklahoma (GRAWIE) 8. Some chemical aspects of the origin of petroleum (LINO) 8. Hydrocarbons (Fr pat 693 054) 10.

Treating cold petroleum oils with sulfuric acid. HORACE B. SETTLER and MERT C. McDONALD (to National Refining Co.) U. S. 1,791,329, Feb 3. The temp of the oil is reduced to about 5° and the oil, during continued agitation, is subjected to the action of H_2SO_4 and clarified oil and sludge are sepd. App is described.

Cracking petroleum oils. DAVID G. BRANDT (to Doherty Research Co.) U. S. 1,791,114 Feb 3. In cracking oil under pressure, charging stock is passed in heat interchange relation to, but out of contact with, vapors in a dephlegmator so that a reflux condensate is produced, this condensate is passed through a heater under pressure to produce highly heated oil and vapors. The highly heated vapors and oil are sepd in a sepg zone, and the highly heated vapors are passed under pressure into a body of oil to crack the latter, residuum is withdrawn from the body of oil last mentioned and is distd and the resulting distillate is independently heated under pressure to a high temp and is passed into the sepg zone. App is described. Cf C A 25, 411.

Cracking crude petroleum. GENERAL TECHNICAL CO., LTD. Fr 693,280, July 4, 1929. The asphaltic residues obtained in cracking crude petroleum are heated with a chem reagent such as baryta and Fe perchloride to about 300° under a pressure below 6 kg., and the products are returned to the cracking operation to be converted into gasoline.

Cracking and destructively hydrogenating hydrocarbons. I. G. PARBENIND A.-G. Brit 335,522 April 23, 1929. The process described in Brit 290,700 (C A 23, 2555) is modified by subjecting the carbonaceous material to a mild cracking process in which extensive formation of benzene is avoided, followed by sepn of any benzene formed and destructive hydrogenation of the remaining material. Catalysts may be employed in both these stages, if catalysts having a splitting action such as bauxite, silica gel, active charcoal or alumina are used for the initial stage. Numerous details and modifications of procedure are described, and an example of the process as applied to gas oil treatment is given.

Cracking hydrocarbons. JEAN P. A. BRUZAC. Fr 693,281, June 22, 1929. Heavy hydrocarbons are converted to lighter hydrocarbons in 2 operations. (1) The hydrocarbons are distd in contact with neutral metals such as Fe, Cu or Ni and condensed by cooling, the temp never being above 500° . (2) The hydrocarbons from the first are submitted to the action of O or ozone to remove excess of C and S.

Cracking hydrocarbons. OWEN D. LUCAS. Fr 692,549, Mar 21, 1930. In converting heavy oils to light oils the temp of the oil to be treated is raised to the final temp required for cracking or conversion by communicating the necessary heat to the oil by indirect contact with biphenyl or phenyl ether in the liquid state in a heat exchanger, this compd being contained in a closed-circuit app to which the heat is furnished apart from the heat exchanger. Before entering the heat exchanger the oil is heated directly to a temp below that at which appreciable cracking takes place.

Hydrocarbons of low molecular weight from those of high molecular weight. N. V. BATAAFSCHE PETROLEUM MAATSCHAPPIJ. Dutch 22,573, Sept 15, 1930. The reaction is accelerated by metal carbides (Ca, U, Mo, Al). The temp is between 300° and 700° , increased pressure may be used. No H_2 , H_2O or H-contg vapors are used.

Cleavage of hydrocarbons boiling up to 40° . N. V. DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ. Dutch 22,574 Sept 15, 1930. Low boiling hydrocarbons propane, butane, etc., are cracked in the presence of a small quantity of catalyst (org or inorg halide, as HCl, CH_3I , CCl_4 , $POCl_3$, etc.) with or without activator (Cu, Fe). Temps of around 600° are used.

Distilling heavy hydrocarbon oils. ROGER D. HUNNEMAN, FRANCIS M. ROGERS and ROBERT E. WILSON (to Standard Oil Co. of Ind.) U. S. 1,791,209, Feb 3. Oil such as residual oil is heated to about $355-405^\circ$ and steam at a like temp in the pro-

portion of 3-10 lb. per gal. of final oil distillate is supplied. Equilibrium is established between the oil and steam while maintaining an absolute pressure not over 75 mm., and the unvaporized oil and vapors are sep'd. App. is described. Cf. C. A. 24, 3333.

Gasoline by fractional condensation. DAVID G. BRANDT (to Doherty Research Co.) U. S. 1,791,113 Feb. 3. Hydrocarbon vapors contg. gasoline constituents, such as vapors produced by distg. and cracking petroleum are passed under superatm. pressure through a series of condensing zones, the last of which is maintained at substantially the normal atm. temp. while preceding zones are maintained at a higher temp., and condensates obtained in the preceding zones are individually cooled to substantially normal atm. temp. and a body of cooled condensate having a higher mean h. p. than that of the condensate obtained in the last condensing zone is maintained and mixed with the condensate from the last condensing zone. An arrangement of app. is described.

Gasoline from fuel oil. L. M. POTIER and F. RASQUIN. Belg. 370,471, June 30, 1930. Fuel oil is treated under pressure in a closed vessel with a compressed gas. The mixt. thus obtained is purified in filtering app. arranged in a closed circuit.

Apparatus for producing gas from gasoline. JAMES T. JONES (to Norman A. Way, Carley Zelmanovitz and Ashton Burford). U. S. 1,791,213, Feb. 3.

Filter for gasoline, etc. AMALGAMATED CARBURETTERS, LTD., and C. BROWN. Brit. 235,691 Aug. 26, 1929. Structural features.

Filter for gasoline, etc. WM. J. VILLAVISO. U. S. 1,791,333, Feb. 3. Structural features.

Filter for oil of internal-combustion engines. ERNEST J. SWEETLAND. U. S. 1,791,010, Feb. 3. Structural features.

Apparatus for dehydrating oil. LEONARD D. GRISBAUM (to Railway Service and Supply Corp.). Can. 308,563 Feb. 10, 1931.

Solutions of lower mercaptan compounds. GERALD L. WENDT (to Standard Oil Co. of Ind.) U. S. 1,791,179, Feb. 3. A soln. of lower mercaptan compds. which is suitable for removing S from gasoline, etc., is prepd. by washing petroleum naphtha with a NaOH soln. and adding to the resulting soln. a Ca compd., such as lime or Ca hydrosulfide, which reacts with the Na sulfide present to ppt. the sulfide radical.

Bituminous emulsions. HERMEL KALTASPHALT A.-G. (C. A. Braun, inventor). Ger. 514,483, Mar. 22, 1925. Stable emulsions of asphalt and other bituminous materials and ceresin are prepd. by supplying the material to be emulsified, together with alk. water contg. less than 1% of alkali, to a mixing vessel contg. a pre-formed emulsion of the material. The amt. of emulsion that can be so prepd. from the batch of pre-formed emulsion is stated to be practically unlimited. Examples are given. Cf. C. A. 24, 4627.

Reclaiming journal box lubricating oil from waste. LEONARD D. GRISBAUM (to Railway Service & Supply Corp.). U. S. 1,791,473, Feb. 3. Contaminated emulsified oil is treated with an alk. soln. such as NaOH of sufficient strength to break the emulsion and react with contaminations including lint, at a temp. above 75° (preferably about 100°), the mixt. is agitated with air, mixed with water, further agitated with air and the oil is sep'd. App. is described. U. S. 1,791,474 describes app. and detail of procedure for treating the contaminated oil with NaOH soln. and projecting water on the surface of the oil after subsidence of the mixt., to travel downwardly. Cf. C. A. 24, 6001.

23-CELLULOSE AND PAPER

CARLETON E. CURRAN

Determination of alpha cellulose. H. PRINCESS. *Kunststoffe* 12, 282-3 (1930).—Cellulose materials analyzed for alpha cellulose by 3 different methods gave different values. The differences in the results are due mainly to differences in the preliminary drying of the sample and to differences in the methods of washing out the alkali and drying the alpha cellulose.

Measuring the viscosity of cellulose and its application to the manufacture of plastics. L. LAISNEY and H. RECLUS. *Rev. gén. mat. plastiques* 7, 3-15 (1931).—After a brief review of previous work on the viscosity of cellulose, a technique for its detn. is described in detail. The method is simple and easy to use in a com. lab., and is sufficiently accurate for ordinary purposes. The cuprammonium soln. is prepd. via Joyner (C. A. 16, 3543) but the NH₄OH is strengthened to about 250 g. NH₃ per l. before soln. of the Cu and again after soln. of the Cu is completed. The cellulose is dissolved in the

presence of H_2 , and the soln. is forced by pressure of H_2 into a H_2 -filled tube, where the viscosity is detd. by the falling sphere method (a 3 mm. sphere falling 50 cm.). The concn. of the soln. should be such as to give a viscosity of 5-100 sec. The viscosity at various concns. was found to be: 0%, 2 sec.; 1%, 10 sec.; 2%, 140 sec.; 3%, 1000 sec. These values practically correspond to the equation $\log \eta = 1.01 C + 0.035$.

A. PAPINEAU-COCTURE

Production of ammonium copper oxide cellulose solutions. E. SCHURZ. *Kunstseide* 12, 283-5(1930). *Jenigen's Rayon Review* 2, 248-9(1930).—A detailed discussion of the prepn. of cellulose cuprammonium solns. involving the following steps: pptn. of $Cu(OH)_2$, mixing of the $Cu(OH)_2$ with cellulose, extn. of the blue liquor, soln. of the $Cu(OH)_2$ cellulose mixt. in ammonia, filtration of the resulting soln., and removal of air bubbles therefrom. The mixing of dry cellulose with solid $Cu(OH)_2$ results in a much better control of the spinning soln.

FREDERICK C. HAHN

Nitrates of cellulose benzoate. M. SENDO AND J. KONDO. *Cellulose Ind. (Tokyo)* 6, 150-5(1930).—Cellulose monobenzoate was treated at 18-20° with mixed acids of the following percentage compn., resp. H_2SO_4 , HNO_3 , HNO_2 , and water: (1) 71.14, 20.81, 0.13, 7.92; (2) 65.6, 18.1, 0.15, 16.15; and (3) 59.58, 21.39, 0.23, 19.8. Cellulose dinitrate mononitrobenzoate $[C_{12}H_9O_5(NO_2)CO_2C_{12}H_7NO_2]$ was obtained by means of acids (1) and (2) and the mononitrate $[C_{12}H_9O_5(NO_2)CO_2C_{12}H_7NO_2]$ by acid (3). The explosive properties of these nitrates are similar to those of the normal cellulose nitrates, but, compared with equally nitrated normal cellulose, these mixed esters are more stable, less hygroscopic and less soluble in ether-alc. and burn at a slower rate.

B. C. A.

Acid adsorption and stability of nitrocellulose. D. R. WIGGAM. *J. Phys. Chem.* 35, 536-9(1931).— HNO_3 and H_2SO_4 are adsorbed by nitrocellulose of high N content; other acids are adsorbed somewhat less. The stability of the nitrocellulose is lowered to a degree depending upon the amt. of acid adsorbed.

MALCOLM DOLLE

Phenomena of the lattice transformations of nitrocellulose. Their generality in cellulose compounds. J. J. TRILLAY. *Compt. rend.* 191, 1441-3(1931).—Trogus, Hess and Katz (*C. A.* 24, 2877) have shown that the x-ray interference pattern of ramie fibers, nitrated to a max., is changed if the fibers are swelled by certain liquids. The altered pattern reverts reversibly to that of the original fiber as the liquid evaporates from the fiber. Cotton, nitrated to 12.95% N (nearly trinitrocellulose), gave a cryst. diagram corresponding to trinitrocellulose. Some of this nitrated cotton was then completely dispersed, instead of merely being swelled, by dissolving it in acetone, and was then recovered in the form of a film by evapn. of the acetone. The films showed a cryst. structure having a new lattice spacing of 9.31 Å. U. and differing slightly in other respects from that of the undispersed cotton. It is uncertain whether this modification of the structure of the lattice corresponds to a new cryst. form of trinitrocellulose, or arises from a combination of the trinitrocellulose with the solvent. The film retains about 1% of acetone very tenaciously. The new diagram appears to be a mixture of that of the original trinitrocellulose with that of a new product. If the degree of nitration of the cotton is less (10-12%), the fibers show a cryst. structure, although more feebly as the degree of nitration is less. The corresponding films, however, show only halos corresponding to an amorphous structure. This difference between the films from cotton nitrated to 12.95% and to 12% is quite sharp. It appears that only trinitrocellulose is a well-defined compd. capable of crystallizing in a second form as the result of passing through the disperse phase. The presence of an ill-defined compd. such as "dinitrocellulose" hinders the process and gives rise to an amorphous structure. Similarly, cellophane is cryst. and has a structure different from that of the initial cotton. In general, it follows that the compd. of cellulose and native cellulose both can assume different cryst. forms after being passed through a state of swelling followed by dispersion.

R. H. LOWHARD

α,β -Lignin. PETER KLASON. *Svensk Kem. Tids.* 42, 259-63(1930).—Reply to Hägglund's contention that there is only one lignin in wood (cf. *C. A.* 24, 5152, 6007).

A. R. ROSE

Utilization of hardwoods for mechanical and chemical pulp. H. S. HILL. *J. Forestry* 28, 1146-52(1930).—As spruce becomes more expensive, hardwoods become more important as pulp raw material. That the pulping qualities of hardwoods have already been investigated and found to be satisfactory for a variety of pulp products is evident from the literature reviewed by H. The future use of hardwoods for pulp is essentially a matter of economics.

ALFRED L. KAMMERER

Wood fibers. G. J. RITTER. *J. Forestry* 28, 633-41(1930), cf. *C. A.* 22, 4149, 23, 976, 1931.—The results of an investigation of the microstructure of the fibers of 10 species of wood are presented. Two types of lignin are distinguished. Fiber cell walls

are made up of layers, capable of being sepd into fibrils by chem and mech means. Fibrils in turn can be chemically sepd into fusiform bodies. Cellulose and lignin from wood can be distinguished by their optical properties in polarized light. A. L. K.

Composition of Cassia slamea, L. Y. SHINODA. *Cellulose Ind* (Tokyo) 6, 155-6 (1930).—The following percentage compn has been found: moisture 11.2, alc benzene ext 18.2, lignin (König) 37.3 (on the extd, dry product), ash 0.3, impure cellulose 40.3, hydrolysis value (Ost) 92.8 and pentosan content (Tollens Kruger) 15.6%. Before chlorination the wood meal was steeped in 10% NaOH soln for 48 hrs and its cellulose content redetd, when the value 33.8%, of purity 100%, was found. Xylose and mannose were found in the hydrolysis liquor, but no galactose was present. B. C. A.

Bamboo. I. Fine structure of the bamboo fiber. K. SISIDO. *Cellulose Ind* (Tokyo) 6, 148-60 (1930).—A theoretical discussion of the results of previous workers. B. C. A.

The use of microorganisms in certain commercial processes. E. E. HUBERT. *J. Forestry* 28, 542-5 (1930).—A résumé is given of making com products with the aid of molds, bacteria yeasts, etc. with bibliography. Expts have shown that spruce and balsam fir attacked by the group of fungi causing white rots may produce relatively large yields of pulp by the sulfite process. This finding has possibilities in the future development of pulp manuf. It is carrying out expts on the action of a blue stain fungus on western yellow pine, the resulting by-product of which has com. possibilities.

ALFRED L. KAMMERER

Removal of hemicelluloses from wood by use of sodium hydroxide. O. HORN. *Cellulosechemie* 11, 151-2 (1930).—Mineral acids and 5% NaOH have been used to remove hemicelluloses, undesirable impurities in wood investigations. In the presence of NaOH, lignin may be oxidized or its components may be split off, perhaps through sapon and the balance may be altered. Beechwood sawdust, after extn with benzene ethanol, 1:1, was degummed with 5% NaOH by 4 treatments at room temp. for 36 and for 48 hrs, the ext. after each treatment being removed and new alkali used. The av. wt. loss for 4 detns. was 23%. If no lignin is removed, a lignin content of 21.1% before degumming must rise to 29.6%, whereas experimentally, it was found to be a little over 27%. Methoxyl content for the wood before degumming was 6.25%, equiv. to 7.75% for degummed wood, whereas experimentally only about 6.3% was found. Each alkali treatment was followed by washing with cold water, dil. AcOH and hot water until the filtrate was colorless, the results in all expts. being fundamentally the same. Conclusion: The method for degumming by use of 5% NaOH is a suitable one because, in addition to hemicelluloses, a part of the lignin is also removed. C. A. BRAUTLICH

Quality of technical caustic soda employed in the manufacture of viscose silk. V. I. SHARKOV. *J. Chem. Ind* (Moscow) 6, 1027-9 (1929).—The purity of NaOH for use in viscose silk manuf. is of the utmost importance. Practical experience has shown that the presence of NaCl + Na₂CO₃ in NaOH in greater concns. than 3.5-4% is harmful, although very low concns. are not markedly bad. These 2 salts exert about an equal depressing effect on the swelling of the fiber of sulfite cellulose. A. C. ZACHIN

Cellulose acetate rayon. K. OGUER. *Kunstseide* 12, 262-7 (1930).—A review of recent scientific works related to the prepn. of cellulose acetate, discussion of these results in the light of present practice in the manuf. of cellulose acetate rayon, and a comparison of this rayon with other types of rayon from various standpoints.

FREDERICK C. HAHN

Dry spinning of acetate rayon. FRITZ OHL. *Kunstseide* 12, 279-81 (1930). *Jentgen's Rayon Review* 2, 245-7 (1930).—Numerous factors influencing the production and quality of acetate rayon are discussed. A cellulose acetate should be chosen with the smallest possible structure of the cellulose molecule; it should be sol. in proper solvents, free from impurities, of uniform viscosity and should not contain over 0.05% H₂SO₄. The spinning soln. must be prepd. under the strictest conditions. A cellulose acetate of fixed moisture content must be used, it should be uniformly dissolved with a min. period of stirring to avoid warming the mixt. and loss of solvent. A no. of filtrations are required and all bubbles must be removed from the soln. prior to spinning. Exptl. data are presented showing the importance of the foregoing factors. In the dry spinning process a careful control of the quantity and temp. of the air is important. F. C. HAHN

Strengthening of rayon. TATSUMI YAMADA. *Repts. Imp. Ind. Research Lab., Osaka, Japan* 11, No. 10 (1930).—Absorption of water by cellulose depends upon the presence of a free OH group and by replacing it with NO₂ or AcO groups it is possible to increase the resistivity toward water. A similar result is obtained by treating rayon with formalin. Y. found the following conditions necessary to treat rayon with formalin:

pr 1.2-2.0, time of immersion 12 hrs., time of heating 5 hrs. at 50-70° and concn. of soln. 4-10%.

F. I. NAKAMURA

Mechanical advances in the artificial silk industry. HANS SCHMIDT. *Chem. App.* 17, 205-7, 220-2 (1930).—A review.

M. C. ROGERS

Beating of [paper] pulp. XI. A characteristic of vegetable fibers as paper-making materials. M. NAKANO. *Cellulose Ind. (Tokyo)* 6, 144-7 (1930); cf. *C. A.* 24, 5491.—Wool silk, asbestos and vegetable fibers were beaten, and their paper forming properties compared. Wool cannot be split longitudinally but is reduced to short lengths showing no fibrillation, from which it is impossible to form a sheet. Silk and asbestos split up into fibrils, but the sheets formed from the products are weak because of the slippage of the fibers. It follows, therefore, that fibrillation is not necessarily accompanied by the development of the adhesive properties which are essential for the production of satisfactory sheets. Only vegetable fibers have these characteristic properties, which are due to the colloidal state of the surface of the fibrils, and their felting capacity is of secondary importance.

B. C. A.

Beating of [paper] pulp. XII. Improvement of absorbency of blotting paper. M. NAKANO. *Cellulose Ind. (Tokyo)* 6, 164-7, *Abstracts* 29 (1930), cf. preceding abstract.—Blotting papers of high absorbency are obtained by the use of pulps treated with concd. NaOH solns., and the absorbing velocity of such products is, on the av., about 40% higher than that of paper made from untreated pulp. Hence pulps can be transformed into "freer" states than their natural ones, and this effect is the reverse of that obtained by beating.

B. C. A.

The utilization of kaoliang stalk. III. Industrial experiments on the manufacture of soda pulp and paper. RYUJI YAMAMOTO. *Abstracts from Rept. Central Lab. S. Manchuria Railway Co.* 1929, 27-7.—70% kaoliang pulp is used to make white or Manila paper.

V. P. HARRINGTON

The Spierer lens and what it reveals in cellulose (SEIFRIZ) 11A. Constitution of cellulose xanthate (LIESER) 10. Lignin and cellulose (FREUDENBERG, et al.) 10. Methylated tri- and tetra-saccharides from cellulose and starch (FREUDENBERG, FREUDENBERG) 10. Concrete grindstones for wood pulp manufacture (Brit. pat. 335,521) 20. Artificial threads, films, etc. (Fr. pat. 633,240) 25. Treating cellulose material (Fr. pat. 632,796) 25.

LEJEUNE-JENO, PAUL. Aus dem Werdegang der deutschen Zellstoff-Industrie 1830-1930. Festschrift zum 50 jährigen Bestehen des Vereins deutscher Zellstoff-Fabrikanten e. V. Berlin Selbstverlag d. Vereins. 112 pp. M. 10

Cellulose. ERIC HIGGINS. Ger 515,018, June 10, 1927. In the manuf. of cellulose by digesting wood with NaOH soln., a part of the black liquor by-product is returned to the digester, without causticizing, after its org. constituents have been destroyed by heating it to about 350° under 20 atm. pressure.

Cellulose. ERNST MEYER. Ger 514,495, Feb. 20, 1929. A pure, highly voluminous cellulose is prepd. by heating com. cellulose for 3-4 hrs. at 3-5 atm. pressure with a dihydric aliphatic alc. or a deriv. thereof, particularly a mono- or di-alkyl ether or an acetate or chloride. A small quantity of a weak alkali, e. g. borax or soap, may also be present. Thus washed sulfite cellulose may be heated as above with ethylene glycol mono-methyl ether 2 and NH₃ 0.5%, and the product washed and bleached.

Cellulose. HANS KAMMERL. Fr 633,005, Mar 29, 1930. A valuable cellulose is obtained from soiled plant fibers, particularly cotton waste such as linters, by treating the fibers mechanically to remove the greater part of the impurities in a centrifuge without a screen, after which the pectins, waxes and greases buried in the structure of the fiber are removed by leaching or light boiling.

Drying cellulose. GUSTAV NEUMANN. Fr 633,014, Mar 29, 1930. Cellulose is dried by means of liquid SO₂, which is then extd. by lowering the pressure. The dried cellulose is particularly suitable for acetylation or to be transformed into viscose.

Acidulation of cellulose. SOC. ANON. FOUR LIND. CHIM. & BAIE. Fr 633,800, Apr 12, 1930. Cellulose is acidulated by treating it first with concd. solns. of neutral salts such as CaCl₂, MgCl₂, ZnCl₂, or thiocyanates of K, Ca, etc., and then with anhydrides of aliphatic acids. Cf. *C. A.* 24, 3112.

Nitrating cellulose, etc. I. G. FARBENFAB. A.-G. (Franz Becker and Hans Zepter, inventors). Ger 515,108, Oct. 30, 1927. A small addn. of urea is made to the nitrating mixt. in order to eliminate or reduce the amt. of N₂O₄. The process is particularly useful in nitrating cellulose with relatively dil. nitrating mixts.

Cellulose esters. I. G. FARBERND, A.-G. (Max Hagedorn and Georg Hingst, inventors) Ger. 515,100, Jan 7, 1927 Sol cellulose esters of higher fatty acids are prep'd by treating unpretreated cellulose with chlorides of the acids in the presence of pyridine, quinoline or like base at a temp above 100°, e. g., 110-140°. Diluents may be present. An example is given.

Cellulose esters. I. G. FARBERND, A.-G. (Max Hagedorn and Georg Hingst, inventors) Ger. 515,107, Jan 29, 1927 Insol or difficultly sol cellulose esters of higher fatty acids are converted into sol esters by treatment in a liquid medium at a raised temp not exceeding 200° with an org or inorg acid, acid anhydride or acid salt, or mixts of these. Thus, insol cellulose tristearate gives a clear soln when heated in $C_2H_5Cl_4$ with Cl_3CCOOH to 145°, and the ester isolated by pouring the soln into $EtOH$, is sol in C_6H_6 , $CHCl_3$ and $AcOAm$. Other examples are given also Cf C A 25, 1379.

Cellulose esters. Soc KODAK-PATHÉ (Soc ANON FRANÇAISE) Fr 693,189, Apr 2, 1930 Cellulose esters of fatty acids are made by heating esterifiable cellulosic material in vapors of a fatty acid having more than one and less than 8C atoms, at a temp above 100° but below 200° in the absence of O and catalysts producing a degradation of the cellulose, until the acyl group in the ester thus produced has reached at least 4%, the said vapors being the sole acylating agent. The ester may be esterified afresh with addnl acyl groups.

Benzylcelluloses. JOSEPH BOUCHET Fr. 693,318, July 12, 1929 Aralkyl-celluloses, and particularly dibenzylcellulose, are obtained by treating alkali cellulose prep'd with concd NaOH lye with $C_6H_5CH_2Cl$ at 70-100° without pressure and with or without a diluent.

Cellulose acetate. OSWALD SILBERRAD and HARRY BLEASHALE Fr 693,568, Apr 9, 1930 The $AcOH$ is recovered from the acetylation mixt by adding to this mixt an amt of water greater than that necessary to destroy all the anhydride remaining, e. g., an addn of 6-8% of water. The mixt. is allowed to stand at 20-35° until the cellulose acetate has become sol in acetone then transferred to a distg app and finely ground $AcONa$ is intimately mixed with it, and the $AcOH$ distd off. Fr 693,569 describes the manuf. of cellulose acetate in which the cellulose acetate is pptd from an acid product and collected as a porous opaque substance by dig the acid product with water in amt. almost sufficient to start the pptn and afterward bringing the acid product thus treated into contact with water so that from the moment of contact the product is submerged in an excess of water.

Cellulose acetate. Soc KODAK-PATHÉ (Soc ANON FRANÇAISE) Fr 693,132, Apr. 1, 1930 Cellulose acetate is prep'd by treating a cellulosic material with Ac_2O and a catalyst in the presence of a mixt which dissolves the cellulose acetate produced, the mixt. contg an amt of $AcOH$ insufficient, by itself, to dissolve the acetate and an org volatile liquid not miscible with water and inert with regard to Ac_2O , e. g., C_2H_5Cl . The cellulose acetate is pptd in an aq bath at a temp above the b. p. of the volatile liquid, which is collected and condensed. Fr 693,133 describes the recovery of cellulose acetate from soln. in $AcOH$ by mixing the soln with an aq pptn bath contg at the beginning at least 20% by wt of $AcOH$ and capable of pptg the acetate.

Cellulose acetate compositions. Soc. KODAK-PATHÉ (Soc. ANON. FRANÇAISE). Fr 693,131, Apr 1, 1930. A compn suitable for making films is made by dissolving cellulose acetate sol. in acetone in a volatile solvent mixt composed of $AcOMe$ 50, $AcOEt$ 10-40 and acetone 40-10 parts. This soln contains about 30% of cellulose acetate.

Apparatus for circulating and storing colloidal solutions of cellulose. Soc. INDUSTRIELLE DE MOIR Fr 692,647, June 24, 1929 The app is constructed so that the cellulose is brought in a perfectly homogeneous state to the spinning nozzles for making artificial silk.

Saccharification of cellulose. MAURICE JUNIEN, Fr 693,277, June 14, 1929. Cellulosic materials are heated under pressure to remove free O and then hydrolyzed with hot dil H_2SO_4 in the presence of excess of water to obtain fermentable sugars. An app. is described with means for keeping the acid in rapid movement during the hydrolysis.

Carbohydrate esters. I. G. FARBERND, A.-G. (Max Hagedorn, inventor). Ger 515,109, Feb 11, 1928 See Fr 668,686 (C A 24, 1650).

Nitrocellulose solutions, etc. JOSEPH G. DAVIDSON (to Carbide and Carbon Chemicals Corp) U. S. 1,791,301, Feb 3 Nitrocellulose or a like cellulose deriv is used with a solvent comprising the acetate of the monoethyl ether of ethylene glycol or other solvent of the same type.

Felted cellulose fiber products. GEORGE L. SCHWARTZ (to E. I. du Pont de Nemours & Co) U. S. 1,791,248, Feb 3 In forming products such as those suitable for

manuf. of "artificial leather," shoe box toes, etc., continuous bands of cellulose fibers are gelatinized with caustic alkali, the gelatinized fibrous material is subjected to heat and pressure and washed. App. is described.

Plastic compositions comprising cellulose esters and polymerized vinyl compounds. CELLULOID CORP. Brit. 335,582, March 23, 1928. Cellulose acetate or other cellulose esters are used with polymerized vinyl compds. such as polymerized vinyl acetate, propionate, phthalate, phosphate, chloride or bromide, styrene or the products obtained by dehalogenating vinyl halides as by use of Zn, with or without various plasticizers or softening agents, fillers, pigments, gums, resins, etc., or anti acid substances such as urea or Ca lactate or volatile solvents.

Plastic masses. I. G. FARBENIND A.-G. Fr. 693,496, April 7, 1930. Plastic masses are composed of derivs. of cellulose such as acetylcellulose in combination with a gelatinizing or plastifying agent formed of one or more esters of phosphoric acid, the org. constituent of which consists of monoalkyl or monoaryl glycolic ethers of the ethyleneglycol, propyleneglycol or butyleneglycol series such as dimethylglycolbutylglycol phosphate or dipropylglycolbutylglycolphosphate. Several examples are given.

Pyroxylin sheets. THE FIREKLOTH CORP. Fr. 693,186, April 2, 1930. Smooth sheets of pyroxylin are made by transforming a gel of pyroxylin into a sheet while still plastic, and modifying the phys. state of the surface of the sheet while drawing. The sheets may be extended and submitted to a tension while extending and a part at least of the solvent is removed by warm air.

Molding celluloid sheets and articles. H. DE HEN. Belg. 360,941, June 30, 1930. Previously heated and softened celluloid sheets are simultaneously pressed and subjected to the action of chemicals and of evapn. until they have acquired the desired shape and such a consistency that they can be removed from the press without being cooled.

Viscose. LEON LILIENTHAL. U. S. 1,790,090, Feb. 3. Cellulosic material such as cotton or wood-cellulose is treated at least once with a caustic alkali soln. contg. not less than 15% caustic alkali (calcd. as NaOH), washed, converted into alkali cellulose, and the alkali cellulose is converted into viscose. Cf. C. A. 24, 2292.

Nontransparent and glossy capsule of regenerated cellulose. KALLE & Co. A.-G. (Julius Voss and Otto Schnecko, inventors). Ger. 515,377, Nov. 7, 1925. See U. S. 1,778,099 (C. A. 24, 6014).

Artificial silk from viscose. ZELLSTOFFFABRIK WALDHOF and A. BERNSTEIN. Brit. 335,605, June 27, 1929. In a noncontinuous process of the type in which the thread, after coagulation and fixing, is collected on spools, and is removed by passing the threads immediately after coagulation through a hot water bath, and the threads are then spooled. The coagulating bath may be formed of water 70, Na₂SO₄ 13, H₂SO₄ 10 and glucose 7%, and formalin may be added to the hot water bath.

Artificial silk from viscose. I. G. FARBENIND A.-G. Brit. 335,675, Nov. 8, 1928. Viscose made from unripened alkali cellulose and contg. excess of alkali as compared with its content of cellulose is spun into a soln. of an NH₄ salt such as (NH₄)₂SO₄ to which acid and other NH₄ salts such as the acetate or formate may also be added, as may also other neutral salts such as Na₂SO₄, NaCl, NaOAc or Na lactate, etc. For subsequent coagulation the formed threads are passed through an acid bath of low concn. such as a 6% H₂SO₄ soln. and the freshly pptd. threads may be subjected to stretching. The viscose is made from unripened cellulose at temps. above 0°. Cf. C. A. 25, 814.

Artificial silk. COMPTON DES TEXTILES ARTIFICIELS. Brit. 335,613, June 28, 1929. In forming filaments such as those from viscose spun into an acid salt bath, the partly coagulated filaments after leaving the coagulation bath are stretched by passage through a narrow tube through which flows a current of liquid (which may be the same as, or different from, the coagulation bath) offering resistance to the forward movement of the filaments as they are drawn by rollers or other devices. App. is described.

Artificial silk. ZELLSTOFFFABRIK WALDHOF and A. BERNSTEIN. Brit. 335,606, June 27, 1929. While passing from a twisting spool to a reel, the material is allowed to fall into a liquid bath preferably contg. a material such as a wax-oil emulsion which forms a protective coating on the filament. The app. is described. Precipitants which do not crystallize out are preferably used in the spinning process.

Artificial silk. BENNO BORZYKOWSKI. Fr. 693,700, April 11, 1930. See Ger. 492,279 (C. A. 25, 415).

Artificial silk. N. V. HOLLANDSCHE KUNSTZIJDE INDUSTRIE. Fr. 693,411, April 5, 1930. Artificial silk of a dull luster is made by introducing into the spinning soln. terpenes or terpenic alcs. or mixts. thereof. An emulsifying agent such as sulfonated castor oil is also added.

Artificial silk. NOVASETA A. G. ARNON. Fr. 691,749, Mar. 11, 1930. Artificial

silk on leaving the spinning nozzles is given a preliminary coagulation in a coagulating bath and then caused to pass in a direction not deviating much from the vertical through at least one subsequent vessel traversed by a pptg liquid to obtain a final coagulation

Apparatus for drawing artificial silk in the moist state. FRANZ J GAULEAT 1r 693,574, April 9, 1930

Twisting and collecting artificial silk, etc. WM. P DREAPER 1 S 1,700,979, Feb 3 Mech features

Artificial threads. MICHEL J DASSONVILLE Fr 692,665, June 26, 1929 Cellulose threads having a structure and qualities comparable to those of natural wool are obtained by varying the section of the thread from the spinning nozzle by a corresponding variation of the flow or of the pressure of the cellulose soln about to be spun

Artificial filaments. ALGEMEENE KUNSTZIJDE UNIE N. V. Ger 514,400, Aug 24, 1927 Artificial filaments of dull luster are prep'd from solns of cellulose acetate or other cellulose ester or ether by dry spinning these in a spinning cell into which water or water vapor is introduced The spinning soln may be anhyd or not Examples are given

Luminous artificial filaments. HENRY A GARDNER U S 1,791,199, Feb 3 Artificial filaments such as those formed from cellulose acetate or viscose contain a comp'd of Ti such as the oxide and a small proportion of a radioactive substance, such as RaBr₂

Removal of hydrogen sulfide from air of viscose-silk factories. N V NEDERLANDSCHE KUNSTZIJDE FABRIEK Dutch 23,233, Dec 15, 1930 The surplus alk liquor from the prep'n of cellulose is used, dil'd to 4%, for absorption of H₂S from the foul air catalyzers such as Ni, Cr and Mn salts (1%), are added for oxidation of the H₂S Cf C A 24,4633

Fiber from wood. GEORGE A RICHTER and MILTON O SCHUR (to Brown Co) U S 1,790,838, Feb 3 Chipped wood such as jack pine is digested under pressure and with heating, in a saponaceous alk liquor such as a NaOH soln contg 0.3-0.4% of soap Cf C A 25,414

Wood pulp. GUSTAF HAGLUND (to Patentaktiebolaget Gröndal Ramen) U. S. 1,791,476, Feb 3 Wood chips, before cooking with a liquor such as a sulfite soln, are preliminarily soaked in a cooking liquor produced in the usual manner without any addn of waste liquors from previous cooking operations, and, after removal of part of the cooking liquor, the chips are cooked with a cooking liquor mixed with waste liquor from a previous cooking operation or with such liquor from a cooking app still in action

Cellulose pulp web. GEORGE A RICHTER (to Brown Co) U S 1,790,839, Feb 3 A web which is highly absorptive of impregnating substances comprises gelatinized cellulose and loosely felted, substantially uncompacted fibers in unbeaten condition

Cooking cellulose. ALB E NIELSEN Norw 45,601, June 17, 1929 The cooking is carried out by indirect steam heating with circulation of the liquor through an app outside the digester During the cooking process the liquor is subjected to a successive evap'n in the circulation app in order to increase the heat value of the waste liquor, reduce the consumption of fuel and raw material and increase the output of pulp

Pulp beater ("pulsing engine"). S MULNE Brit 335,671, July 26, 1929

Apparatus for continuous beating and brushing of paper-making stock. THOMAS J. MULLEN U S 1,790,830, Feb 3 Structural features

Washing paper-making pulp. OTTO MANTJUS U S 1,790,714, Feb 3 App and various details of app and operations are described

Paper. FAIENRICH K WICKEL Ger 515,186, Mar 27, 1928 Impervious transparent paper is prep'd from sulfite cellulose paper by impregnating this, as completely as possible, with anhyd glycerol, and then coating it on both sides with an aq soln of albumin, preferably egg albumin, to which addns such as hardening agents may be made A soft and non adhesive product is obtained after drying Cf C A 24,1741

Unbleached paper, cardboard, etc. VALDEMAR W D'OBRY. U. S. 1,791,092, Feb 3 A grass is moistened and steamed under pressure while preventing escape of substantially all the constituents of the grass, and the material is then mech disintegrated for making paper, cardboard or the like Cf C A 24,1951

Paper-making apparatus. DEE L SHEAFOR (to Black Clawson Co). U. S. 1,791,384, Feb 3 Structural features

Paper-making apparatus. NILS A JACOBSEN U. S. 1,791,412, Feb 3 A device is provided, actuated by the wt of torn off paper, for automatically dividing the paper web on the wire cloth in 2 or more parallel strips in case tearing of the paper web should be started because part of the paper is carried along with the wire cloth about the couch roller

Suction box for paper-making apparatus. EARL E. BERRY (to Beloit Iron Works). U. S. 1,700,852, Feb. 3. Suction boxes are made of comparatively thin wall material composed of a Cr alloy such as "stainless steel" having a high modulus of elasticity and resistant to corrosion by the liquids with which it comes into contact.

Waterproofing paper. WM H. RICHARDSON (to Richardson Co.). U. S. 1,791,040, Feb. 3. In making waterproof cartons, a blank is printed and sprayed with paraffin on the printed side, and afterward is calendered with calender rolls. App. is described.

Composition for waterproofing paper, millboard, papier maché articles, etc. DUNLOP RUBBER CO., LTD., D. F. TRISS and G. GORHAM. Brit. 335 559, June 25, 1929. A material suitable for treating containers for food consists of an aq. emulsion or dispersion of rubber, gutta percha, balata or the like, in admixt. with an aq. dispersion of a mineral or vegetable wax or mixt. of waxes, e. g., paraffin, in which the wax content may amount to 20-75% of the "dry rubber wax content." The compn. may also be used for sealing the lids of the containers.

Dyeing paper and paper pulp. SOC. ANON. FOUR L'IND. CHIM. A. BALE. Fr. 693,172, April 2, 1930. A unilaterally mono- or dialcylated diamine such as oleyldiethylethylene-diamine is used as adjuvant in dyeing paper or paper pulp.

Recovery of paste from printed paper, and chiefly wood paste. LOUIS GRENAUDIER. U. S. 1,791 445, Feb. 3. Materials such as old newspapers are soaked in water and then treated with an alk. bath such as Na_2CO_3 to form an alk. soap with the fatty matter of the ink, and a substance such as NaCl is used which renders the soap insol. so that the ink-contg. soap rises to the top of the bath and does not mix with the paste in the bottom of the bath.

24—EXPLOSIVES AND EXPLOSIONS

CHARLES B. MCVOR AND C. G. STORM

Drying powder and explosive materials in vacuum. I. BODEWIG. *Chem. App.* 17, 253-4, 279-80 (1930).—B. reviews the difficulties encountered in drying powder and tells of improvements in the app. M. C. ROGERS.

High brisance studies on explosive effect and chemical constitution. I. Tetranitromethane and its mixtures with toluene. ALFRED STETTBACHER. *Z. ges. Schiess-Sprengstoffw.* 25, 439-41 (1930).—S. describes the prepn. of $\text{C}(\text{NO}_2)_4$ from $(\text{C}_2\text{H}_5\text{CO})_2\text{O}$ and HNO_3 in lab. app. Because of the time required for the reaction (10-12 days) and poor yield (56%) the product has no industrial application. Its properties are described. When properly neutralized and dried it is apparently quite stable. A mixt. of 100 parts $\text{C}(\text{NO}_2)_4$ and 15.66 parts $\text{C}_{11}\text{H}_{11}\text{Cl}_3$ has a heat of explosion of 1701 Cal/kg. The sensitiveness of this mixt. is greater than that of nitroglycerin, and its brisance as indicated by Fe plate tests is even greater than that of gelatin penthrinit (C. A. 24, 6017, 25, 818). C. G. STORM.

High brisance studies. II. Mannitol hexanitrate (nitromannitol) and pentaerythritol tetranitrate (Penthril). ALFRED STETTBACHER. *Z. ges. Schiess-Sprengstoffw.* 25, 461-2 (1930).—In comparative tests of nitromannitol and "Penthril," alone and in mixts. S. again attributes the greater explosive strength of the "Penthril" to its mol. structure about a central C atom (C. A. 24, 6017). Approx. 30 g. samples were detonated in heavy Fe crucibles on brass or Fe plates, these being preferred to Pb plates because of their greater hardness. Among other mixts. tested, one contg. nitromannitol 37.5%, nitroglycerin 56.2% TNT 6.3% was compared with "Penthril" 34%, nitroglycerin 51%, nitrocellulose 2.15%, NH_4NO_3 12.9%, both mixts. having an exact O-balance. The latter mixt. gave slightly superior results. A frozen mixt. of 60% "Penthril" and 40% nitroglycerin (60/40 "Penthrinit") is claimed to have the highest rate of detonation of known explosives—9500 m./sec., as against 9000 m./sec. for frozen nitroglycerin. C. G. STORM.

The preparation of nitrocellulose with maximum nitrogen content. F. LENZE AND E. RUMENS. *Z. ges. Schiess-Sprengstoffw.* 26, 4 (1931).—Nitrating expts. were made with cotton linters and mixts. of HNO_3 with H_2SO_4 , SO_3 , $(\text{CH}_3\text{CO})_2\text{O}$, N_2O_5 and P_2O_5 in varying proportions. A max. N content of approx. 14% was obtained with a nitrating mixt. of HNO_3 and P_2O_5 . The product was only slightly sol. in Et_2O : EtOH mixt. Further expts. are in progress. C. G. STORM.

The properties of the nitroglycerin isomers. A. A. DERSHEVICH and K. K. ANDREEV. *Z. ges. Schiess-Sprengstoffw.* 25, 353-6, 400-3 (1930).—The literature of the subject is reviewed. D and A compared liquid nitroglycerin (A) with the frozen labile

(B) and stable (C) forms, as regards brisance and rate of detonation. The Pb block test method of Hesse showed (C) to have much greater brisance than either (A) or (B), and also to be more sensitive to detonation. The same relation was confirmed by the Kast method (compression of Cu cylinders), although this method gave discordant results. Rate of detonation tests were made by the Dautriche method in glass tubes 22 mm inside diam. with 2 mm walls, picric acid detonating fuse having a rate of 7000 m/sec was used. (C) gave the remarkably high rate of 9150 m/sec (av. of 3), (A) gave 1165 m/sec. (B) failed to detonate completely with No. 8 detonators, but gave a rate of 9100 m/sec when a booster of 20 g tetryl was used. Similarly, (A) gave a rate of 8750 m/sec when a booster of 10-15 g (C) was used. Contrary to the rate of detonation results (B) gave the highest Traut block test (500 cc), (A), 519 cc, (C), 390 cc. In this test the insensitiveness of (B) was probably overcome by the confinement afforded by the sand stemming. No appreciable difference was found as regards sensitiveness to impact. When properly protected from contamination, (B) remained unchanged for periods as long as 70 days. The linear conversion velocity and crystal velocity of the isomers were determined. It is suggested that the loss of sensitiveness of gelatin dynamites during storage is the result of gradual conversion of (A) to (B).

C. G. STORM

Researches on coal-mining explosives at the Safety in Mines Research Station, Burton. W. PAYMAN *Trans Inst Mining Eng* 80, 11-23 (1930-1931).—Tests are made to determine the safety of explosives in dusty and gassy coal mines. These tests include the number of cartridges in a shot-hole, the position of the detonator, the character of the stemming, the safety in the presence of coal dust and of methane. Ignitions of gas have been obtained with as little as 2 oz. of a permitted explosive with no stemming, but 1 in. of noncombustible stemming will prevent ignition of gas with 28 oz. of a similar explosive. The presence of an inhibitor to prevent ignition of gas by explosives has been considered but this method has not been found practical. Expts on the effect of compression of gas have indicated that, up to at least 11 atm., as the pressure of the air and methane is increased the ignition point is lowered. When the pressure is reduced the ignition point of methane and air is reduced even more rapidly than under increased pressure. A max temp is reached in the case of methane under about $1/2$ of an atm. The greatest amt. of work is done when a detonator is placed in the cartridge next to the stemming and pointing inward.

C. W. OWINGS

Explosive limits of fire damp. E. KIRST *Glückauf* 67, 50-7 (1931).—Results of recent investigations were discussed.

J. D. DAVIS

A review of some recent dust explosions. D. J. PRICE AND H. R. BROWN *Quart Natl Fire Protect Assoc* 24, 305-29 (1930).—A comprehensive report is given on dust explosions at Western Maryland Grain Elevator, Hogan Mill Feed Co., L. F. Carlston Elevator Co., Tobacco Byproducts and Chem Corp. and Staley Manufacturing Co. All these dust explosions occurred during August and September, 1930, resulting in a gross loss of approx \$500,000. Analysis of all that is known of the causes further emphasizes the necessity of cleanliness in industries producing flammable dusts, importance of controlling equipment design to minimize the likelihood of stirring up dusts or igniting them, and the value of inert gases as a method of preventing dust explosions in grinding equipment.

C. L. JONES

Nitrates of cellulose benzoate (SENDO, KONDO) 23. Acid adsorption and stability of nitrocellulose (WIGGAM) 23.

SACHSE, FRICH. Die Bekämpfung der Schlagwetter- und Kohlenstaubgefahr. Berlin: Phoenix-Verlag C. Schwinn. 32 pp. M 2.40

Explosives. DYNAMIT A.-G. VORM ALFRED NOBEL & CO. (Pb Naoim and H. Ulrich, inventors) Ger 513,653, Mar 28, 1930. Added to 500,407. The invention of Ger. 500,407 (C. A. 24, 4397) is extended to cover the nitrates of homologs and of N-alkyl derivs of monoethanolamine and diethanolamine. Cf C A 25, 1383.

Gunpowder blasting aquibs for electrical ignition. IMPERIAL CHEMICAL INDUSTRIES, LTD., and W. MCGHIE. Brit. 335,607, June 27, 1929. Structural features

25—DYES AND TEXTILE CHEMISTRY

L. A. OLNEY

Standard colors and fewer dyes would benefit textile industry. ALBIS MAERZ, *Textile World* 79, 428-30 (1931)

Color measuring and color standards. M. J. SCHOEN *Chem Weekblad* 28, 105-6 (1931)

Photoelectric cell has definite place in color measurement. WALTER M. SCOTT *Textile World* 79, 895-9 (1931) —A résumé is given of the accomplishments to date along the line of color measurements, together with a warning to textile manufacturers "that they must not look for an immediate millennium in this field." The applicability, limitations and future possibilities of the photoelectric cell for color measurement are discussed. The following color analyzers that use this device are briefly described: the T. C. B. Photocolorimeter, designed by Toussaint, the photoelectric spectrophotometer developed by the American Photoelec. Corp. marketed by Eimer & Amend, the portable recording and indicating color analyzer developed by P. J. Mulder and J. Rawek, marketed by the Burgess-Patt Co., the Colorscope, developed by H. H. Sheldon, New York Univ., and the recording photoelectric color analyzer developed by A. C. Hardy, manufactured by General Elve Co. The fallacies of the following claims are pointed out: that the photoelectric cell introduces a distinctly new method of color measurement, that it entirely eliminates the human eye and therefore greatly increases the accuracy of the measurement by doing away with the variable human factor, and that photoelectric instruments are so simple to operate that no previous specialized training is necessary. R. K. W.

Relation between structure and affinity of dyes for plant fibers. MICHAEL DOKSIEWICZ, *Przemysł Chem* 14, 229-48 (1930) —See C. A. 24, 5501. A. C. Z.

The vat colors from a cotton piece dyers' point of view. GEO. H. HARRIS, *Dyer, Calico Printer* 64, 643-4 (1930) —A brief discussion based on practical experience is given of the difficulties involved in the use of vat colors. The selection of a suitable dyestuff, and the importance of such factors as the quality of the water, the temp. of the dye bath, the method of handling the goods in running them through the dye bath, etc., are considered. F. K. W.

Pectographic studies of solutions of dyestuffs. PAUL BARY, *Rev. gén. colloïdes* 8, 62-8 (1930), *C. A.* 24, 2417 —The similarly constituted compounds pararosaniline and fuchsin AB in aq. solns. give the same sort of pectographs in dil. and concd. solns., with crystals being retarded in dil. solns. Methyl violet does not give crystals. Safranin T gives crystals at all concns., while safranin MN gives amorphous bands but yields crystals at 0.025% concn. Hence, in safranines, dila. aids crystal. and impedes it in the triphenylmethane derivs. The choice of solvent modifies completely the pectographs obtained in almost all cases. Picric acid in acetone gives clear unique crystals while in alc. it gives a gel impregnated with crystals. Safranin MN, which in water gives an unstable gel and crystals, in acetone yields fine lines as also do safranin T, methyl violet, pararosaniline, crystal violet and induline scarlet. Ultramicroscopic exams. of these solns. show that all those which give a pectograph of fine parallel lines contain visible granules while those which give a gel deposit are optically void. P. W. L.

Action of nitrous acid on animal fibers, and reactions of the products formed with azo components. M. J. VAN TUNENBERG, *Chem Weekblad* 28, 62-5 (1931) —The question is reviewed whether HNO_2 forms reactive diazonium compds. with amino acids of the fiber, or whether it is merely absorbed, leading, on development, to nitrosophenols and their dye derivs. Formation of $\text{HO}(\text{C}_6\text{H}_4)_2\text{N}_2\text{O}$ from PhOH and N_2O_5 in excess (Morel and Sailey *C. A.* 22, 4515) was confirmed in this particular instance only. When N_2O_5 was conducted into an EtOH soln. of *p*-cresol and tyrosine, $\text{HO}(\text{O}_2\text{N})_2\text{C}_6\text{H}_3\text{Me}$ was formed, but no diazo compd. Light fastness of dyes developed on the fiber is very good, especially upon afterchroming. Shades are decidedly changed (deepened) and acid and alkali fastness improved by chroming. Diazotized fibers show light sensitivity. Photographs may be printed out on silk. K. H. ENGEL.

Rapid identification of dyes on fibers. RAYMOND LEVÉ, *Téla* 9, 11-7 (1931) —A scheme of analysis is presented for the rapid identification of the class of dye used on a given sample. The various tests (treatment with HClO_4 , boiling with Na_2CO_3 , reduction with SnCl_2 and with alk. hypophosphite) are carried out by adding the reagents to the dyed fibers in test tubes. A. PAPINEAU-COUTURE.

The dyeing of cotton warp during sizing. A. POUTET, *Téla* 8, 1447-9 (1930) —The dyeing of cotton warp in the sizing bath and before sizing is described and a no. of formulas for each process are given. A. PAPINEAU-COUTURE.

Chemical finishing treatments of cotton. C. THUNAU *Textile World* 8, 1457-63 (10-10), 9, 35-41 (10-11). The treatments, other than mercerization and bleaching, which effect a considerable modification in the phys. aspect of cotton are described. A. P.-C.

Dyeing and finishing cotton corduroys. L. MACKINNA *Textile Recorder* 48, No. 475, 57 (10-11). A detailed description is given of the processing of this type of fabric from the time it comes from the loom in the grey to the final folding and boxing of the finished article. RUBY K. WORKMAN

Continuous bleaching challenges the chemist. WINN W. CHASE *Textile World* 70, 616-7 (10-11). The need of chem. developments to keep pace with the mech. improvements is stressed. RUBY K. WORKMAN

Practical suggestions for better handling of acetate yarns. EDWARD H. GILBERT *Textile World* 70, 141 (10-11). Suggestions are given for modifications and adjustments of the machinery used in winding, twisting and tinting cellulose acetate yarns. RUBY K. WORKMAN

Preparation of fabrics made of, or containing, acetate rayon, prior to dyeing. J. CALABIN *Kunstf.* 5, 1871, 1873 (10-10). A. PAINHAU COURTINE

Dyeing of acetate rayon. FREDERICK BRANDENBURGER *Kunststoffe* 12, 267-72 (10-10) cf. C. A. 25, 704. Photochemical decomposition of the dyes. The fading of dyes on acetate rayon fabrics is discussed in the light of oxidation, reduction and enzyme theories. Most of the cases discussed are explained on the basis of oxidation. According to this theory the dyestuff forms a peroxide which oxidizes the constituents of the dye. The dyeing of cellulose acetate is viewed as a soln. process and not as an absorption process. The dye is present in a more finely divided state than on cotton textiles and therefore a greater surface of the dye is exposed to oxidation. On this basis is explained the poor fastness of basic dyes and achl. (wood) dyes on acetate rayon. The fastness of various diam. colors and its relation to their constitution are discussed. *Ibid.* 305-10. A retrospect and an outlook in the future of acetate rayon development with regard to the dyeing of this type of rayon and its mixts. with cotton, viscose rayon, wool and silk. FREDERICK C. HAHN

Dyeing acetate rayon black. JEAN PAUL SIEURY *Textile World* 9, 17-23 (10-11).—A review of the various dyes and processes which have been proposed. A. PAINHAU COURTINE

Practical viscose dyeing. HANS ENGEL *Kunststoffe* 12, 200-2 (10-10); *Senigen's Rayon Review* 2, 255 (10-10).—Whitely dyed viscose rayon is as a rule caused by insufficient preliminary washing and improper handling during dyeing, and, less frequently, by the choice of dyestuff. Streakiness in the dyeing is due to abnormal viscose. The foregoing difficulty may be alleviated by the use of dyes of the Rigan type, or mixts. thereof with other direct dyes which are known to equilibrate well. F. C. H.

The structure of rayon in relation to anilino-black dyeing. A. J. HATT *Silk J.* 7, No. 77, 37-8, No. 78, 33, 11 (10-10), cf. C. A. 24, 6021.—Because of the greater absorbent properties of the regenerated rayons they give a better black than cotton. Acetate rayon shows a relative absorbency for basic compds., and will not absorb salts. Aniline and other amines are absorbed but the oxidant (NaClO_2) is not. It is suggested that acetate rayons might be treated with aniline first and then oxidized in a fresh bath. At present aniline black has not been satisfactorily dyed on acetate rayons. A. R. MACGORMAC

Preparation, dyeing and finishing of rayon crepes. A. CRISP *Textile World* 5, 60-63 (10-11). A. PAINHAU COURTINE

Printing of fast-color rayon crepes requires careful attention to details. HARRISON S. STRICKLAND *Textile World* 70, 574 (10-11). RUBY K. WORKMAN

Sizing and finishing of rayon. W. A. DYER *Kunststoffe* 12, 313-6, 310-2 (10-10); cf. C. A. 24, 1086.—A discussion of the prepn., testing and required properties of starch pastes for the sizing and finishing of rayon. FREDERICK C. HAHN

The coloration of silk stockings. S. R. TROTMAN *Dyer, Calico Printer* 65, 23-4 (10-11); cf. C. A. 21, 1085. The various processes are described with particular emphasis on the difficulties involved. The water used, the preliminary soaking, the app. used for dyeing, the degumming process, the dyeing of pure silk and silk cotton stockings, the selection of suitable dyestuffs for each type, etc., are discussed. RUBY K. WORKMAN

Woolery redyeing made easier. FREDERICK C. ROBERTS *Textile World* 70, 316 (10-11).—Before stippling a preliminary bath composed of 1-3% soda ash and 1 1/2-1 3/4% mol. pine oil is recommended. The advantages of the use of the mol. pine oil are: It penetrates and opens up the fibres without injuring them. It removes part of the color, but does not start the resist colors in the pleat edge or check. It reduces the amt. of hyposulphite or sulfoxylate needed. It reduces the stippling time, thereby reducing also the danger of mech. injury to the goods. RUBY K. WORKMAN

The preparation of tussah silk and schappe silk prior to dyeing. J BRACONNOT. *Russa* 5, 1867-71, 2043-7(1930), 6, 41-5(1931) —A review A P C

The printing of wool and silk using the Neolan colors. VERNON D FREEDLAND *Dyer, Calico Printer* 64, 646-7(1930) —The Neolan dyes used for printing are divided into those which require the addn. of Cr acetate to the printing paste and those which do not require it Recipes are given for each type A list of the solubilities of these dyestuffs in H_2O at $212^{\circ}F$ is also given RUBY K. WORTNER

Plate printing on silk. J CLASOV *Russa* 5, 1705-9(1930) —Practical operating hints, with a no. of formulas A PAFINEAU-COUTURE

The dyeing and finishing of union wool-silk goods. L BONNET *Russa* 5, 2047-57(1930) —This is a general description, dealing more particularly with the various types of machines available for the handling of the goods A PAFINEAU-COUTURE

Dyeing wool with indigo in fermentation vats E DUTOIT *Tiba* 8, 1309-19(1930) —The principles and technic of the process are described A P C

The treatment of lace curtains. C L FRISWICK *Tiba* 8, 1435-41(1930) —Methods of bleaching, dyeing and finishing of lace curtains, at present in use in various European countries, are described A PAFINEAU-COUTURE

Dyeing and finishing curl cloths. DOUGLAS KERMODE *Dyer, Calico Printer* 65, 139-40(1931), cf *C A* 24, 6022 —Curl cloths consist of a worsted or worsted mohair pile in loop form, knitted into a cotton backing As a rule its manu. is a work for specialists Dyestuffs which have been found suitable for this type of fabric are listed and methods of applying them are described RUBY K. WORTNER

The dyeing of wool-felt and hair-felt hats. E DUTOIT *Tiba* 8, 1441-7(1930) —Practical operating hints with a no. of formulas are given A PAFINEAU-COUTURE

Dyeing vegetable-ivory buttons KARL HAHN *Dyer, Calico Printer* 64, 399-400(1931). —See *C A* 24, 5502 RUBY K. WORTNER

Bleaching and dyeing leaves and grasses. F GROVE-PALMER *Dyer, Calico Printer* 64, 707-8(1930), cf *C A* 25, 561 —The basic dyes are generally used for these materials because of their brightness, cheapness and high tinctorial power Cellulose enamels in various colors may be sprayed on with an air pistol Metal leaves can be obtained by spraying on bronze powder in a clear cellulose medium, or better still by electroplating with metals RUBY K. WORTNER

Kaoliang starch as a finishing material for textiles. JUSABURO KAWAMATA *Abstracts from Rept. Central Lab. S. Manchuria Railway Co.* 1929, 27-9 —In respect to its value as a sizing material kaoliang starch occupies a position between wheat starch and corn starch V F HARRINGTON

Noteworthy machinery development of recent years marches unretarded through 1930 into 1931. EDWIN D FOWLE *Textile World* 79, 604-7(1931) —A review of new machinery developed for bleaching, dyeing and finishing RUBY K. WORTNER

A new use for textiles. GEO RICE *Dyer, Calico Printer* 65, 28-9(1931) —Cotton stationery is described with particular reference to the processes used for stiffening, coating, weighting and glossing the fabric. The requirements of this material are that its stiffness shall be such that it can be written upon with a pen or printed upon with a typewriter, that the interstices shall be filled, and the finish smooth, so that the impression of the pen or the typewriter will be even that the handle shall be satisfactory, that the writing shall be retained even when wet and that the fabric must permit erasures RUBY K. WORTNER

Washing cotton and linen. YNGVE DALSTRÖM *Svensk Kem. Tids.* 42, 263-77(1930) —A crit. review is given of the conflicting laundry technology literature A R ROSE

Future textile-laboratory practice. GEO B HAVEN *Am. Dyestuff Rept.* 19, 757-61(1930), *Textile World* 78, 3028-30(1930), 79, 42-4(1931), *Melland* 2, 1319-28(1931) —New avenues are suggested for directing the future work of the Textiles Committee of the Am. Soc. for Testing Materials Illustrated descriptions are given of app. for measuring yarn balance, corkscrew in yarns, crimp and yarn slippage, all of which are properties of the yarn, and for measuring porosity, absorbing power for liquids, bursting strength, precise thickness, abrasion, heat flow and resilience of fabrics R K W

Influence of laundering and exposure to light upon some washable silks. MARION GRIFFITH *Ohio Agr. Expt. Sta., Bimonthly Bull.* 147, 179-81(1930) —Only pure dye branded silks were used All 12 samples showed a very decided fading upon exposure to light and during the laundering process Blue and green colors faded most, while peach and yellow were least affected by light exposure The highest priced silk contained less than $\frac{1}{2}$ the amt. of finishing materials present in the least expensive silk The breaking strength varied from 31-69% of the original when exposed 48 hrs. to light The de-

crease in strength resulting from 15 launderings varied from 15-17%. The silks lost 25-42% of their original elasticity as a result of exposure to light for 48 hrs. The expts. are being continued along chem lines. C. R. FELLERS.

A cloth-wear testing machine. J. A. MATTHEW *J. Textile Inst.* 21, 540-60T (1930).—The design of machines for testing cloth wear is discussed and a new machine is described. Its principle is that a sample under a definite tension is rubbed with a Carborundum surface. The no. of rubs necessary to produce fracture is taken as a measure of the resistance of the cloth to wear. Results are given to show the effect of varying the rubbing conditions for a cloth and also to show the comparative resistance to wear of a no. of cloths under fixed rubbing conditions. It is shown that wear tests and tensile strength tests bear little relation to each other. RUBY K. WORKER.

Relation of cotton to synthetic fibers. C. E. MULLIN *Jenigen's Rayon Review* 2, 231 8(1930).—See *C. A.* 24, 5483. FREDERICK C. HAIN.

Fiber cross-section. H. L. BARTHELEMY *Kunstseide* 12, 316-7(1930).—An improved method of making rayon cross sections is described. A slit is made lengthwise in a small cork to a depth $\frac{1}{4}$ of the diam. A single thread of rayon is placed in the slit parallel to the axis of the cork and then disks are cut therefrom with a razor. This method is an advantage over the Herzog method in that the individual filaments can be distinguished and counted. FREDERICK C. HAIN.

Fiber cross-section of acetate rayon. H. STANGLER *Jenigen's Rayon Review* 2, 240-3(1930), *Kunstseide* 12, 310-3(1930).—The two general types of acetate rayon cross sections, "ribbon like" (I) and "cloverleaf" (II) are discussed, and microphotographs are given of the cross sections of various com. acetate rayons. Earlier acetate rayons were of the type I but present-day rayons are of type II. Type II are superior as regard to lower gloss and higher wet and dry strengths. FREDERICK C. HAIN.

Physicochemical studies on the structure of wool hair after treatment with alkali, acids and chlorine. ERNST TAYLER. *Wiss. Arch. Landw., Abt. B., Tiererndhr. Tierzucht* 4, 297-346(1930).—Wool from several breeds of sheep was treated with HCl, H_2SO_4 , HNO_3 , $AcOH$, Cl water and chloral hydrate. The wool was then investigated as to histological changes, change in wt., fineness, flattening, hygroscopic properties, tensile strength, elasticity, polarization, reaction with methylene blue, biuret test, reaction with Millon's reagent and Allwörden's reaction. Many references are appended. W. GORDON ROSE.

Rotenone as a moth-proofing agent. E. A. BACK, R. T. COTTON AND R. C. ROARK. *J. Econ. Entomol.* 23, 1014(1930).—Rotenone, the insecticidal constituent of derris root, cubé root and some other tropical plants is effective as a moth proofing agent. From 1 to 2% of rotenone is dissolved in acetone and the soln. is used to impregnate woolen goods. The impregnated goods appear to be highly resistant to attack of the clothes moth, *Tineola bisselliella*, and of the beetles, *Anthrenus rorax* and *Altigenus piceus*. Acetone solns. contg. as low as 0.05% rotenone gave excellent protection from these insects. The rotenone treatment appears to be equal to that of other proprietary moth-proofing solns. now on the market. Application for a public service patent has been filed. C. H. RICHARDSON.

Note on a little-known fireproofing process. E. DUTOIT *Tida* 9, 33(1931).—Perkin's method consists in impregnating flannel with a 26° Bé Na stannate soln., draining, drying between heated Cu rolls, treating with a 10° Bé $(NH_4)_2SO_4$ soln. to ppt. $Sn(OH)_4$, draining, drying, washing and drying again. The resultant fireproofing is permanent, as the SnO_2 is combined with the fiber, it increases the mech. strength of the fabric and its resistance to washing and ironing, and moreover it acts as mordant if it is to be dyed. A. PAPINEAU COUTURE.

Condensation products from arylthioglycolic acids [dyes] (GEBAUER-FOLNFGG JARSCH) 10. Heteropolar C compounds XII. New dyestuffs of the aniline blue series and perchlorates of several important triphenylmethane dyes (DILTHEY, DIK-LAGE) 10. β -Hydroxy- α -picoline azo dye (Brit. pat. 335,818) 10. Leuco indigo (Fr. pat. 693,469) 10. Triazine derivatives [intermediate] (Brit. 335,783) 10.

Dyes. COMPAGNIE NATIONALE DE MATIÈRES COLORANTES ET MANUFACTURES DE PRODUITS CHIM. DU NORD RÉUNIES (ÉTABLISSEMENTS KUHLMANN) Fr. 692,927, June 23, 1929. Anthraquinone derivs. which dye cellulose acetate silk are prep'd by the reaction of sulfonic acids of aminohydroxyanthraquinones such as diaminoanthranufin and diaminochryzazine, their salts or reduction products, with either an aldehyde or its substitutes in the presence of hyposulfite, or hyposulfite compds. of aldehydes previ-

ously formed in the presence of an excess of hyposulfite or sulfoxylates of aldehydes. Thus, Na diaminoanthraquinone 2,6-disulfonate is dissolved in water and BzH and $\text{Na}_2\text{S}_2\text{O}_4$ are added and the mixt. is heated. The product dyes cellulose acetate blue. Several other examples are given.

Dyes. I G FARRENTD A-G. Fr 693,474, April 7, 1930. Dyes of the anthraquinone series are made by the reaction of α -aminoanthrimsdicarbazoles or their substitution products with anthraquinone- α -halocarboxylic acids or their derivs. or substitution products and transforming the reaction products into the corresponding acridones by the known methods. Thus, 4-amino-1,1'-anthrimsdicarbazole (by sapon of 4 benzyl-amino-1,1'-anthrimsdicarbazole) is heated with anthraquinone 2-bromo-1-carboxylic acid in the presence of AcOK , MgO and $(\text{AcO})_2\text{Cu}$. The product dyes cotton from the vat in olive green shades. Other examples and formulas of the products obtained are given.

Dyes. I G FARRENTD A-G. Fr 693,109, April 2, 1930. *N* condensation products which dye cotton from the vat are prepd. by condensing aminoanthraquinone-carboxylic acids, the COOH groups of which are so transformed that they do not form salts, with negatively substituted polynuclear org. compds. Thus, dibromo-3,4,8,9-dibenzopyrene 5,10-quinone is boiled in C_6H_6 with NaOAc , CuO and the Et ester of 1-aminoanthraquinone 2-carboxylic acid until practically all the Br is removed. The product dyes cotton in violet red shades from the vat. Several other examples are given.

Dyes. I G FARRENTD A-G (Paul Ochwat, inventor). Ger. 515,331, May 2, 1929. The dyes obtainable from anthrimides by acid or alk. condensation are halogenated, and the products are caused to react with *N* compds. contg. at least one exchangeable H atom. Products giving deeper shades than the initial dyes are so obtained. Thus, the dibromo deriv. of a dye prepd. as described in Ger. 507,340 (C A 25, 599) may be boiled for 2 days in C_6H_6 with 1-aminoanthraquinone in the presence of Na_2CO_3 , NaOAc and CuCl . Other examples are given also.

Dyes. I G FARRENTD A-G (Paul Nawiasky, inventor). Ger. 515,096, July 17, 1928. Addn to 493,232. The method of purifying *N*-dihydro-1,2,2',1'-anthraquinonarine described in Ger. 493,292 (C A 24, 3907), is now applied to the dye obtained by treating the above compd. with H_2SO_4 of low water content, preferably in the presence of H_3BO_3 . An example is given. Cf C A 25, 599.

Azo dyes. ERNEST F GEETHA (to Dow Chemical Co.). U. S. 1,790,807, Feb. 3. Azo dyes which give various red shades are formed from the tetrazo deriv. of a diamino-diaryl ether such as diaminodiphenyl ether coupled with an arylide of 2,3-hydroxynaphthoic acid such as 2,3-hydroxynaphthanilide.

Azo dyes. GUILLAUME DE MONTMOLLIN (to Soc. anon. pour l'ind. chim. à Bâle). U. S. 1,791,444, Feb. 3. By coupling arylides of 2,3-hydroxynaphthoic acid such as the arylide derived from α -aminodiphenyl ether with diazo compds. such as those derived from 4-chloro-2-aminodiphenyl ether, α -aminodiphenyl ether, the corresponding cresyl-chlorophenyl or chlorocresyl ether or the benzyl ethers of α -aminophenols or α -aminocresols, fast red dyes are produced, which may be formed on a suitable substratum and employed for making red varnishes. Examples and details of procedure are given.

Azo dyes. MAX SCHMID and FRITZ STRAUß (to Soc. anon. pour l'ind. chim. à Bâle). U. S. 1,791,472, Feb. 3. Azo dyes suitable for conversion into metal derivs. such as those of Cr or of Cu are obtained by coupling a diazo-compd. contg. an OH group in *o*-position to the diazo group (such as the diazo deriv. of 4-chloro-2-aminophenol-6-sulfonic acid or 1-diazo-2-naphthol-4-sulfonic acid) with pyrazolone derivs. such as 5-pyrazolone-3-carboxylic acid or its Me or Et ester. Examples are given of dyes the Cr compds. of which produce violet or red dyeings.

Azo dyes. I G FARRENTD A-G. Brit. 335,555, June 21, 1929. Dyes insol. in water are formed in substance, on the fiber or on a substratum by coupling the diazo compd. of an amine of the general formula $\text{aryl NH R' X R'' NH}_2$ (in which R' and R'' are aromatic nuclei which may be substituted or unsubstituted and X represents either a direct linkage or an $-\text{N}-\text{N}-$ group, or in which $-\text{R}'-\text{X}-\text{R}''$ stands for a condensed aromatic system) with an arylide of 2,3-hydroxynaphthoic acid. By use of 2 mol. proportions of nitrite, nitrosodazo compds. are obtained and the corresponding nitrosodazo dyes may be treated in substance or on the fiber with a reducing or saponifying agent to remove the nitroso group. Several examples are given. Cf C A 25, 1092.

Azo dyes. I G FARRENTD A-G (Rudolf Heil, inventor). Ger. 515,230, Sept. 9, 1928. Brown azo dyes insol. in water are prepd. by coupling the diazo compds. from 2-amino-*N*-alkylpyrazoleanthrones, or 2-amino-*N*-aralkylpyrazoleanthrones, with arylides of 2,3-hydroxynaphthoic acid. An example is given.

Azo dyes. IMPERIAL CHEMICAL INDUSTRIES, LTD Fr. 692,694, March 3, 1930 See Brit 331,839 (C A 25, 213).

Monoazo dyes and amines. BRITISH RESEARCH ASSOCIATION FOR THE WOOLEN & WORSTED INDUSTRIES AND A T KRUG Brit 335,646, July 6, 1929 Monoazo dyes and amines are produced simultaneously by the reduction of disazo dyes of the type $XN_2NYN_2C_6H_4(OH)$ in which X and Y represent aromatic residues such as those of benzene, naphthalene, diphenyl, diphenyl sulfide or other grouping the reduction being effected with alkali metal bisulfite (with intermediate formation of disazo-sulfites) The disazosulfites may be treated with alkali metal hyposulfite to produce the amine or aminoazo compd Several examples are given

Disazo dyes. I G FARBEIND A G Brit 335,705, Oct. 4, 1928 A disazo dye giving vivid greenish yellow dyeings on wool, fast to felling and to light, is formed by coupling tetrazotized 4,4'-diamino-3,3'-dichloro-5,5'-dimethyltrphenylmethane with 1 (2'-chloro-5'-sulfo)phenyl 3-methyl-5-pyrazolone The 4,4'-diamino-3,3'-dichloro-5,5'-dimethyltrphenylmethane is made by condensing BzH with 3-chloro-2 toluidine Mention is also made of the production of some other similar dyes Cf C A 24, 2304

Disazo dyes. I G FARBEIND A G (Karl Dobmaier, inventor) Ger 515,231, Aug 21, 1927 Disazo dyes giving yellow shades on cotton are prepd by coupling the diazo compds from nitrobenzoylated diamines or their derivs with suitable components free from NH₂ groups, reducing the NO₂ group of the monoazo dyes so obtained, diazotizing, and coupling with pyrazolone derivs contg diazotizable NH₂ groups, or with compds. of the C₆H₅ or C₆H₄ series in which a diazotizable NH₂ group is present in a "heteronuclear" side chain Suitable components for the first coupling are phenols and their ethers and carboxylic acids, acetoacetanilide, and phenyl methyl pyrazolone For the second coupling, m aminobenzoyl 2,5,7 aminonaphtholsulfonic acid is specified, *inter alia* The reduced monoazo dyes may be condensed with nitrobenzoyl chloride or its derivs., reduced again, and diazotized before finally coupling as above Other modifications are described also Examples are given

Disazo and trisazo dyes and their metal compounds. SOC. ANON POUR L'IND CHIM À BAIE, F. STRAUB AND W ANDERAU Brit 335,523, April 23, 1929 Disazo and trisazo dyes obtainable by coupling with the monoazo dye from 2-amino-5-naphthol 7-sulfonic acid (2 mols coupled acid) either 1 mol proportion of an o hydroxydiazocompd or 2 mol proportions of an o-carboxy diazo compd, or one mol proportion each of an o-hydroxydiazocompd and a diazo compd contg no o hydroxyazo group (excluding in this case o-hydroxycarboxy-group contg diazo compds) are treated (at any stage) with chroming or other metallizing compds such as compds of Mn, Fe, Co or Ni Various examples are given for producing dyes of different colors.

Trisazo dyes. IMPERIAL CHEMICAL INDUSTRIES, LTD Fr 693,585 April 9, 1930 Trisazo dyes, contg at least 2 COOH or 2 SO₃H groups or one of each and suitable for dyeing regenerated cellulose, are obtained by coupling tetrazotized 3-aminobenzene-1'-azo-4' aminonaphthalene or a substitution product thereof with one mol. proportion of a phenol naphthol or a carboxylic or sulfonic acid thereof and with one mol proportion of the same or another component other than a naphtholarylketone or an arylide of 2,3 hydroxynaphthoic acid The same products are obtainable by coupling diazotized 3-nitrobenzene 1'-azo-4'-aminonaphthalene or a substitution product thereof with a phenol, naphthol or a carboxylic or sulfonic acid thereof, reducing, diazotizing and coupling with a suitable component Several examples are given

o-Carboxyazo dyes. I G FARBEIND A G Fr 693,024, Mar. 29, 1930 o-Carboxyazo dyes dyeing substantively and contg Cu are made by the reaction of agents liberating Cu on dyestuffs obtained by coupling tetrazotized 4,4'-diaminodiphenyl-3,3'-dicarboxylic acid with 2 identical or different mols of arylamino- or aroylaminonaphtholsulfonic acids which contain in the aryl or aroyl group groups which render the compd sol in water Several examples are given

Vat dyes. I G FARBEIND A-G (Arthur Wolfram and Emil Hausdörfer, inventors) Ger. 514,432, Dec. 31, 1927. See Brit. 303,375 (C A 23, 4579).

Vat dyes. I. G. FARBEIND A-G (Paul Nawiasky and Alfred Ehrhardt, inventors) Ger 514,433, July 20, 1929 Green vat dyes are prepd by halogenating dimethoxydibenzanthrone in an inert org solvent in the presence of a catalyst, e. g. FeCl₃ Examples are given

Vat dyes. I. G. FARBEIND A-G (Max A Kunz and Karl Köberle, inventors). Ger 515,328, Mar 31, 1929 Addn to 499,169 (C A 24, 4168) and 510,600 (C. A. 25, 1094) Violet to green-blue vat dyes are prepd by treating dibenzanthrone or isodibenzanthrone or their derivs with Br or other brominating agent in an acid medium in the presence of a metal or a metal compd. as a catalyst. Hg, Mn and Sb are suitable

catalysts, and when using these the use of HClSO_4 as the acid medium does not give rise to products contg Cl. The products are fast to light and atm influences. Examples are given.

Vat dyes. I G FARBENIND A-G (Robert Berliner, inventor) Ger. 515,329, July 5, 1928. See Brit 314,809 (C. A. 24, 1520).

Vat dyes. I G FARBENIND A-G (Max A. Kunz and Karl Köberle, inventors) Ger. 515,330, Dec. 30, 1928. Addn to 513,228 (C. A. 25, 1396). New vat dyes giving varied shades are prepd by condensing neg substituted dibenzanthrones or isodibenzanthrones with amides or imides of mononuclear carboxylic acids. The condensation may be effected in a high boiling solvent in the presence of a catalyst and an acid binding substance. The products may be saponified by the usual methods to yield mono- or poly-amino compds of the dibenzanthrone or isodibenzanthrone series. 1 example is given.

Vat dyes. I G FARBENIND A-G Fr. 692,869, Mar. 28, 1930. Halogenated benzanthronepyrazolcanthrone, which dye textile fibers in valuable blue shades, are prepd by treating benzanthronepyrazolcanthrone, its derivs of isomers, preferably in the presence of catalysts, with halogens or substances capable of liberating halogens in the presence of diluents or solvents. Several examples are given.

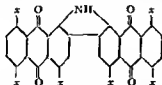
Vat dyes. I G FARBENIND A-G Fr. 693,042, Mar. 29, 1930. Stable leuco vat dye preps forming by soln in water a vat ready to be used for dyeing, are prepd by mixing with a protective colloid, a wetting agent and a stabilizer, the leuco compds of vat dyes of the anthraquinone and thiondigo series which are capable of dyeing from a "cold vat" of hyposulfite. The mixt is dried and the dry powders ground with a caustic alkali and a reducing agent. Thus the leuco compd may be mixed with molasses, an alkyl naphthalenesulfonic acid, hydroquinone, caustic alkali and $\text{Na}_2\text{S}_2\text{O}_4$.

Vat dyes. I G FARBENIND A-G Fr. 693,107, April 1, 1930. Vat dyes contg one or more halogens are prepd by treating isodibenzanthrone or its derivs, or derivs of dibenzanthrone, with anhyd. metallic halides capable of yielding halogens. If necessary other halogenation agents, substances lowering the m. p., appropriate catalysts, solvents or diluent are added. The halides of Fe, Al, Sb and Hg are suitable for the chlorination, and alkali and alk. earth salts may be used for lowering the m. p.

Vat dyes. I G FARBENIND A-G Fr. 693,410, April 5, 1930. Vat dyes are prepd by heating 2 mercapto-1-methylantraquinone or 3 mercapto-2-methylantraquinone or their derivs with S or by heating 2 halo-1-methylantraquinone or 3-halo-2-methylantraquinone or their derivs with S and substances producing or facilitating the substitution of halogens by S. Several examples are given.

Vat dyes. I G FARBENIND A-G Fr. 693,817, April 12, 1930. Vat dyes are obtained by treating the anthraquinonyl-4'-amino-1,2-benzanthraquinones with condensing agents. Thus α -anthraquinonyl-4'-amino-1,2-benzanthraquinone when heated with moderately dil. H_2SO_4 or with a mixt of NaCl and AlCl_3 gives a product which dyes pure yellow from the vat. The same initial material treated with AlCl_3 in the presence of a tertiary base such as pyridine gives a greenish blue from the vat. Several other examples are given.

Vat dyes (anthraquinonylcarbazoles). FRITZ BAUMANN (to General Aniline Works) U. S. 1,791,184, Feb. 3. Vat dyes of the general formula.

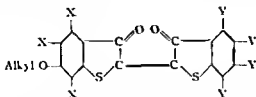


wherein at least one x stands for an anthraquinonylamino residue being attached to the nucleus by only one linkage in the position x, and the other x's stand for substituents of the group consisting of H and any univalent substituent, and wherein all nuclei may be further substituted generally red to grayish powders, insol in org solvents sol in concd H_2SO_4 with a blue to reddish brown color and dyeing cotton from the vat red to grayish black shades of excellent fastness, are produced by the reaction of a 2,2'-dianthraquinonyl 1,1'-carbazole such as 5,5'-diamino-2,2'-dianthraquinonyl 1,1'-carbazole, with the corresponding 4,4'-diamino compd or 5-amino-4'-benzoylamino-2,2'-anthraquinonyl 1,1'-carbazole, with a halogen anthraquinone compd such as 1 p tolylamino-3-bromoanthraquinone, β -chloroanthraquinone or 1-benzoylamino-4-chloroanthraquinone, at an elevated temp (suitably 180-250°) for 3-4 hrs (suitably in a high boiling solvent

such as naphthalene, trichlorobenzene or ethylcarbazole and in the presence of an acid-binding substance) Cf C A 25, 1306

Vat dyes. KARL WILKE (to General Aniline Works) U S 1,700,780, Feb 3. See C A 24, 2305, 2611

Substituted thiolindigo vat dyes. KARL SCHIRMACHER, KARL ZAIEN, ERWIN HOPPA and HANS HEYNA (to General Aniline Works) U S 1,700,843, Feb 3 Dyes of the general formula



wherein X and Y represent H or any substituent, at least one Y being a substituent (the constitution being such that the left hand side of the formula is not identical with the right hand side of the formula) are in a dry state red powders, insol in water, alc. and ether, sol in nitrobenzene with a red color, sol in concd H_2SO_4 with a green color, forming with an alk. hyposulfite yellow vats, from which animal and vegetable fibers are dyed red shades fast to light, Cl and washing, and are obtainable either by subjecting the 1-anils of such diketodihydrothionaphthenes as contain an alkoxy group in *m* position to the S atom, and also their further substitution products substituted in the nucleus, to reaction with substituted ketothionaphthenes, the corresponding arylthioglycolic carboxylic acids or with the acylketothionaphthenes obtainable from the said acids by condensation with acetic anhydride, or inversely by subjecting 5-alkylhydroxythionaphthenes, their substitution products substituted in the nucleus or the corresponding arylthioglycolic carboxylic acids or acylketothionaphthenes obtainable from the said acids by condensation with acetic anhydride, to reaction with the 1-anils of substituted diketodihydrothionaphthenes. By introducing further substituents into one or the other nucleus it is possible to influence the shade toward blue or toward yellow. Numerous examples with details are given.

Stable water-soluble vat dye preparations. JAMES MORTON, JAMES I. M. JONES, BIRKETT WYLAN and JOHN E. G. HARRIS (to Scottish Dyes, Ltd.), U. S. 1,700,750, Feb 3. Vat dyes (numerous examples of which are given in detailed examples of treatment) are treated with reagents such as chlorosulfonic acid, its salts, fuming H_2SO_4 , SO_3 or salts of pyrosulfuric acid, in the presence of a tertiary org. base such as pyridine and of Cu or Zn (suitably with addn. also of a small proportion of $SnCl_4$). The treatment is applicable to dyes of the flavanthrone, indanthrone, dibenzanthrone, benzanthrone, anthraquinone, indigo and halogenated indigo groups, the products being suitable for dyeing cotton, wool or natural or artificial silk from aq. soln. with development of the color of the original dye on treatment with $FeCl_3$ soln., alkali metal hyposulfite or a perborate or permanganate soln.

Indigoid vat dyes. I. G. FARBERIND A-G. Fr 693,903, April 14, 1930. 6-Methylhydroxythionaphthene is condensed with cyclic diketones of the isatinic series or their derivs. Thus, isatin is converted into the α -chloride by means of PCl_5 with the addn. of $PhCl$, 6-methylhydroxythionaphthene in $PhCl$ is added at about 80° and the mixt. stirred. The product dyes cotton in intense bluish grey shades. Other examples are given.

Indigoid dyes. I. G. FARBERIND A-G (Wilhelm Bauer, inventor) Ger 513,230, June 13, 1928. Violet indigoid dyes are prepd by coupling 5-halo-3-hydroxythionaphthenes with reactive α derivs. of haloisatins, especially of 5,7-dichloro- or -dibromo-isatin. An example is given.

Indigoid dyes. I. G. FARBERIND A-G (Karl Thiess, Theodor Meissner and Hans Heyna, inventors) Ger 515,134, Feb 8, 1928. 4,7,4',7'-Tetramethyl 5,5'-dihaloidindigos are prepd by the customary processes from 1,4-dimethyl-3-amino-6-halobenzenes, or from 1,4-dimethyl-3-amino-6-halobenzene-2-carboxylic acids. The same products may be obtained from 1,4-dimethyl-3-aminobenzene or 1,4-dimethyl-3-aminobenzene-2-carboxylic acid, by converting these into 4,7,4',7'-tetramethylindigo and then halogenating. Examples are given.

Indanthrone dyes. LYNNE H. ULICH (to Newport Chemical Corp.) U. S. 1,700,887, Feb 3. See Brit. 314,803 (C. A. 24, 1520).

Wool dyes. CHRMISCHP FABRIK VORN SANDOZ Ger. 515,055, July 13, 1929.

Reddish blue to reddish violet dyes for wool, etc., are prepd by condensing anthraquinone derivs contg substituents exchangeable for an arylamino group with aromatic amines having at least one halogen atom in the *m*-position to the NH₂ group. The condensation products may be sulfonated. Thus, 1-amino-2-methyl-4-bromoanthraquinone may be condensed with *m*-chloroaniline at 160°, and the product sulfonated. Other examples are given also.

Acid wool dyes. I. G. FARBERND A.-G. (Georg Kalischer and Ernst Honold, inventors). Ger 514,434, Feb 2, 1928. See Fr 668,871 (C. A. 24, 1747).

Carboxylic acids of *m*-hydroxyphenylarylamines. I. G. FARBERND A.-G. (Leopold Laska and Oskar Haller, inventors). Ger 515,208, Dec. 2, 1928. Alkali salts of *m*-hydroxyphenylarylamines are heated with CO₂ under pressure. Thus, the K compd of *m*-hydroxydiphenylamine, heated with CO₂ at 150-70° under 5-10 atm pressure, yields a product *m*. 180-1° believed to be 3-hydroxydiphenylamine-4-carboxylic acid. Numerous compds derived from substituted *m*-hydroxydiphenylamines are described also. The products are stable and are useful as intermediates for dyes.

Dibromodibenzanthrone. MAX A. KUNZ and KARL KÖBERLE (to General Aniline Works). U. S. 1,791,215, Feb 3. By a process involving brominating dibenzanthrone with Br in chlorosulfonic acid with addition of Sb, there is obtained a substantially pure dibromodibenzanthrone which dissolves in 96% H₂SO₄ to a pure violet soln, in PhNO₂ to a pure blue soln with a pure red fluorescence and in PhNH₂ to a pure blue soln without fluorescence, and which dyes cotton from a blue vat with a violet tinge strong marine blue shades which are faster to light than the dyeings obtained from the unhalogenated initial material.

Pyrazolone derivatives. I. G. FARBERND A.-G. (Hans Grotowsky, inventor). Ger 514,421, July 17, 1928. Colored condensation products useful as dyes or as intermediates for dyes or drugs are prepd by treating 1-phenyl-3-methyl-5-pyrazolone, or its derivs substituted in the phenyl residue, with benzalaniline or other benzalbenzalamines. The reaction is effected in aq acid soln. at atm. temp. The products are believed to be benzalpyrazolones. Examples are given.

Dye intermediates. BRITISH RESEARCH ASSOCIATION FOR THE WOOLEN & WORSTED INDUSTRIES and A. T. KING. Brit 333,645, July 6, 1929. Aromatic amino derivs are prepd by reduction of monoazo dyes other than those contg a β naphthol component or which form azo-sulfites or which are reducible by SO₂ by the action of alkali hyposulfite and sulfite approx. in the mol ratio of 2NaHSO₃:Na₂SO₃. Examples are given of the reduction of benzeneazosalicylic acid with an alic. soln. of the mixed reagent to form aniline and aminosalicylic acid and of the reduction of Na azosalicylic acid to aminosalicylic acid.

Intermediates for dyes. I. G. FARBERND A.-G. (Winfrid Hentrich, Josef Hilger and Rudolf Knoche, inventors). Ger 515,468, Feb 19, 1929. β -Naphthylaminophenoxy fatty acids are prepd by the reaction of β -naphthol or β -naphthylamine or their substitution products with sulfites and 3- or 4-aminophenoxy fatty acids or their derivs, together or in turn. Thus, Na 2-hydroxynaphthalene-3-carboxylate, heated in aq soln with Na 4-aminophenoxyacetate and NaHSO₃, loses CO₂ and yields 4- β -naphthylaminophenoxyacetic acid, *m*. 156°, after acidifying and cooling. Numerous other examples are given also.

Photochemically decolorizable dyes. I. G. FARBERND A.-G. Fr 693,360, April 4, 1930. Layers for photochem decoloration are prepd by using dyes the speeds of decoloration of which have been harmonized or regulated by hydrogenation or substitution in the aromatic groups joined directly or indirectly to the chromogen. Hydrogenation or the introduction of halogens or a cyano, thiocyan, NO₂, carboxyl or a carbalkoxy group increases the speed of decoloration, while the introduction of an alkyl, aryl, OH, aryloxy or alkoxy group decreases the speed of decoloration. Several examples are given.

Naphthalene derivatives. I. G. FARBERND A.-G. Fr 692,716, Mar 25, 1930. Derivs of C₁₀H₈ are made by the reaction of carboxylic acid chlorides by the Friedel and Crafts reaction on 1-benzyl-naphthalene or on 1-alkyl-naphthalenes or their substitution products which are not substituted in the 4 position of the C₁₀H₇ nucleus, and treating, if necessary, the resulting products with oxidizing agents. The products are intermediates for dyes. Examples are given of the prepn of 4-benzoyl-1-benzyl-naphthalene, *m*. 110°, 1,4-dibenzoylnaphthalene, *m*. 106°, 4-benzoyl-1-(*p*-chlorobenzyl)naphthalene, *m*. 123°, 4'-chloro-1,4-dibenzoylnaphthalene, *m*. 118°, 1,4-di-*p*-chlorobenzoylnaphthalene, and 4-acetyl-1-benzyl-naphthalene, *m*. 78°, also 4-methyl-1-naphthyl phenyl ketone, 4-benzoyl-1-naphthoic acid, 4-methyl-1-naphthyl methyl ketone and 4-acetyl-1-naphthoic acid, *m*. 184°.

Dyeing. I. G. FARBENIND A-G. Fr. 693,426, April 5, 1930. Very clear colored effects are obtained by printing on a support, previously impregnated with arylides of 2,3 hydroxynaphthoic acid and dried, with printing pastes composed of Na formaldehyde hyposulfite and vat dyes, drying again, vaporizing rapidly, and finally coloring with a diazotized base. Examples are given.

Applying dyes to fabric. WILLIAM WHITTENRAD (to Camille Dreyfus). Can. 308,605, Feb. 10, 1931. Dyes are applied to a fabric contg cellulose acetate by spraying with a material contg a water insol dye and ethylene dichloride and a thickening agent.

Dyeing fibrous material. I. G. FARBENIND A-G. (Paul Rabe, Hermann Stötter, Berthold Wenk, and Wilhelm Schepss, inventors). Ger. 515,032, Dec. 23, 1925. In fixing dyes on fibrous materials by means of complex compds of tungstic or molybdic acid these compds are treated on the fiber, before dyeing, with reducing agents, e. g., NaHSO₃, Na₂S-O, or CH₃O.

Dyeing cellulose esters and ethers. I. G. FARBENIND A-G. (Hermann Wagner, Adolf Kuchenbecker, Richard Huss and C. Erich Müller, inventors). Ger. 511,205, Mar. 4, 1928. Addn to 509,401 (C. A. 25, 421). Fast blue dyeings are produced on cellulose esters and ethers by coupling with 2,3 hydroxynaphthoic acid or its arylides the diazo compds from dialkyl, diaryl, or diaralkyl ethers of 2-amino-5-acylaminohydroquinones or their substitution products. Coupling is effected on the material to be dyed. Examples are given.

Dyeing cellulose esters and ethers. I. G. FARBENIND A-G. (Wilhelm Eckert and Carl L. Müller, inventors). Ger. 515,029, Jan. 26, 1928. Cellulose esters and ethers are dyed in yellow shades by means of 4-aminonaphthalimide or its derivs in which an alkyl, aryl, or aralkyl group replaces H in the imino group. Examples are given.

Dyeing viscose products. I. G. FARBENIND A-G. (Karl Bormann, inventor). Ger. 515,030, Feb. 15, 1928. Uniform dyeings are obtained with dyes prepd. (1) by coupling the tetrazo compds of *m* and *p* azoaniline and their substitution products with suitable components, or (2) by coupling the diazo compds. from 3- and 4-nitroaniline or their substitution products with suitable components and reducing the products with alk. reducing agents. Numerous examples are given.

Dyeing artificial silk. IMPERIAL CHEMICAL INDUSTRIES, LTD. Fr. 693,179, April 2, 1930. Regenerated cellulose silk is dyed in relatively regular shades with dyes obtained by tetrazotizing a mono- or disulfonic acid of 4,4'-diaminodiphenylurea or one of its radical derivs and coupling with 2 mol. proportions of a sulfonic acid or carboxylic acid deriv of a phenol, naphthol, naphthylamine or *N*-substituted naphthylamine, or with one mol. proportion of a coupling constituent, as above mentioned, and one mol. proportion of an aminonaphtholsulfonic acid or one of its *N*-substituted products or one of its derivs. Examples and a list of components with the shades obtained are given.

Dyeing artificial silk. IMPERIAL CHEMICAL INDUSTRIES, LTD. Fr. 693,790, April 12, 1930. Silk made from regenerated cellulose is dyed by the products obtained by combining a diazotized amine of the C₆H₅ series, contg no NO₂ group, with an aminonaphtholsulfonic acid in alk. soln., diazotizing again and coupling with another coupling component, all the aminonaphtholsulfonic acids used as final constituents being also coupled in alk. soln. An example and a table of components with the shades obtained on viscose silk are given.

Dyeing furs, hair, feathers, etc. I. G. FARBENIND A-G. (Paul Virck, inventor). Ger. 515,031, Sept. 15, 1926. The materials are treated with an aq. soln. of a salt of a naphthalene deriv. contg at least one 3-aminoaryl amino group of the C₆H₅ series. They are then treated with an acidified nitrite soln. Thus, fur is colored yellow by treatment, at atm. temp., in a bath contg 2-(3'-aminophenylamino)naphthalene HCl, followed by treatment with NaNO₂ soln. acidified with AcOH. Cf. C. A. 25, 605.

Coloring textiles. HENRY DREYFUS. Fr. 692,734, Mar. 25, 1930. In producing discharge effects on textiles, the material is impregnated with a substance having an acid reaction such as an org. acid before applying the discharge agent. The discharged materials may be colored with acylaminoanthraquinone dyes.

Printing with vat dyes. CHARLES J. SALA (to E. I. du Pont de Nemours & Co.). U. S. 1,790,950, Feb. 3. As dye assistants in calico printing pastes, etc., for aiding transfer of the dye to the fabric and producing improved results, there are added substances of the general formula R₁R₂NAikOH in which Aik represents an aliphatic radical as, for instance, the alkylene group C₂H₄, R₁ represents either H or a hydroxylaliphatic group, as CH₂CH₂OH and R₂, as R₁, represents H or a hydroxylaliphatic group. The use of mono-, di- and triethylamines is particularly mentioned, also amines contg al-

phatic groups with more than one OH group such as a compd. with the group $\text{CH}_2\text{CH}(\text{OH})\text{CH}_2\text{OH}$. An example with details of procedure is given.

Machine for dyeing and treating textiles. SMITH, DRUM & Co. Fr. 693,201, April 2, 1930

Machine for dyeing barks of thread. GAETANO S. RICCARDI. Fr. 693,037, Mar. 29, 1930

Condensation products. I. G. FARREBUND A-G. Fr. 692,890, Mar. 28, 1930
New products which are of technical value, e. g., *reserving agents for wool or silk*, are prepd. by condensing a halide of an aromatic carboxylic or sulfonic acid with an aromatic diamine to produce deriva. of the carboxylic or sulfonic acid amide sol. in water which contain no free NO_2 , NH_2 or OH groups. At least one of the constituents of the reaction is substituted in the position *m*- to the halide or to one of the amino groups. The reaction takes place in water or an org. solvent in the presence of an agent fixing hydrogen halide. Thus, bis(2,4-dichlorobenzoyl)-1,3-phenylenediamine-4-sulfonic acid is prepd. by condensing 1,3-phenylenediamine-4-sulfonic acid with 2,4-dichlorobenzoyl chloride and bis(3,4-dimethylphenylsulfonoyl)benzene-*m,m'*-disulfonic acid from 3,4-dimethylbenzenesulfonic acid and benzene-*m,m'*-disulfonic acid. Several other examples of compounds of this nature are given.

Reserve effects on textiles. I. G. FARREBUND A-G. Fr. 692,888, Mar. 28, 1930
Reserve effects are obtained on half-wick half wool fabrics by adding to the dye bath contg. a substantive dye or to the mordanting bath, amides of sulfonic or carboxylic acids in which the H atoms of the amide groups are wholly or partly replaced by alkyl, aryl or aralkyl groups. Examples and a long list of substances which may be used are given, including bis(3,4-dichlorophenylsulfonoyl)benzene-*m,m'*-disulfonic acid (by condensing 3,4-dichlorobenzoylsulfonoyl chloride with benzene-*m,m'*-disulfonic acid in the presence of an acid neutralizing agent) and bis(3,4-dichlorophenylsulfonoyl)-3'-amino-phenylsulfonoyl-3"-aminophenylsulfonoylbenzene-*m,m'*-disulfonic acid (by the reaction of *m*-nitrophenylsulfonoyl chloride with benzene-*m,m'*-disulfonic acid, reducing to the diamine, condensing with *m*-nitrophenylsulfonoyl chloride, reducing again and condensing the product with 3,4-dichlorobenzoylsulfonoyl chloride).

Textiles. HERMANN BOLEMAN and BRUNO REWALD. Fr. 693,887, April 14, 1930
Textiles are improved by adding to the baths used for treating fibers, threads or cloth a certain amt. of leather.

Waterproofing textiles. ALEXANDER NATHANSON. Fr. 693,803, April 12, 1930
Textiles are made waterproof by incorporating in them, by a gentle treatment with esterifying deriva. of higher fatty acids, the corresponding radicals of fatty acids in amts. which do not notably pass the extent of the natural fibers in these radicals. Examples are given of the treatment of textiles with stearyl anhydride, stearyl chloride, etc. in benzene, C_6H_6 , CCl_4 , etc.

Sizing or softening agent for fibers. HERMANN BOLEMAN and BRUNO REWALD. Fr. 692,603, Mar. 21, 1930
A sizing, gumming or softening agent for fibers is composed of an emulsion of phosphatides of vegetable origin (particularly phosphatides from soy beans), fatty oils and (or) fatty acids in water. A small quantity of Turkey-red oil or mineral oil may be added to the emulsion.

Improving the qualities of fibrous materials. I. G. FARREBUND A-G. Fr. 693,019, Mar. 29, 1930
Fibrous materials such as artificial silk are improved by applying to their surface or incorporating with them aliphatic or cycloaliphatic compds. of high mol. wt. or mixed aliphatic or cycloaliphatic-aromatic compds. which possess aliphatic or cycloaliphatic radicals of high mol. wt. or their salts.

Treatment of fabric. HENRY PLATT and CYRIL M. CROSS (to Camille Dreyfus). Can. 208,437 Feb. 3, 1931
Fabric contg. cellulose acetate has the "sticking" point increased by treatment for 45 min. at 40° with a soln. contg. 10-50 g. of basic Al acetate per l. and about 50-100 cc. of AcOH per l.

Treating cellulose materials. HENRY DREYFUS. Fr. 692,796, Mar. 26, 1930
Materials of or contg. cellulose have their qualities improved by a treatment with organo-mineral acids or concd. aq. solns. thereof e. g., a'phatic or aromatic sulfonic acids. The threads may be spun directly from the spinning nozzles into the acids. Examples are given of the use of sulfosuccinic acid.

Artificial threads, films, etc. HENRY DREYFUS. Fr. 693,240, April 3, 1930
Threads, films, ribbons, etc., made of or contg. cellulose deriva. are treated with organo-mineral acids such as sulfonic acids or concd. solns. thereof whereby the mech. resistance and elasticity are improved. Solns. of the cellulose deriva. may be extruded into the acids. Examples are given of the use of sulfosuccinic acid, ethylene-sulfonic acid and methylsulfonic acid. C. I. C. A. 25, 1193

Artificial silk. BRITISH CELANESF, LTD., and G. H. ELLIS Brit 335,583, March 16, 1929. Materials of cellulose acetate or other cellulose derivs. rendered insensitive to the delustering action of moist steam or other hot or boiling aq. media as described in Brit 332,187 (C. A. 25, 205) and Brit 332,231 (C. A. 25, 218) are used together with other cellulose ester or other products sensitive to such delustering action, so that by use of a delustering material on the product design or pattern effects may be produced by its selective action. Numerous details and examples are given.

Artificial wool. MICHEL DASSONVILLE Fr 603,321, July 12, 1929. A machine is described for treating cellulosic filaments coming from the coagulating baths to give them the appearance of wool.

Oleaginous composition suitable for conditioning wool. AUGUSTUS H. GILL, U. S. 1,791,037, Feb. 3. In order to prevent spontaneous combustion of material to which oleaginous compds. are applied, about 1-2% of hydroquinone is added to the compn., e. g., to "red oil" for lubricating wool.

Washing wool, textiles, etc. I. G. FARBENIND A.-G. Fr 692,834, Mar. 27, 1930. Aromatic sulfonic acids or their salts, possessing in the ring or in a substituent or both, one or more org. radicals with a chain of more than 5 C atoms, are used for washing wool, textiles, etc. Examples are given of the use of the Na salt of hexylnaphthalene, octylnaphthalene, cetylnaphthalene- and *N*-dihexylnaphthylaminesulfonic acid. Substances such as cyclohexanol, AmOH, cyclohexanone and CCl_4 may also be added. Cf. C. A. 24, 734.

Mothproofing wool, etc. I. G. FARBENIND A.-G. Brit. 334,886, June 10, 1929. Materials such as woollen fabrics are treated with compds. obtained by condensing phenols with aralkyl compds. such as aralkyl halides, alics., and ω -sulfonic acids, and then sulfonating, or by condensing sulfonated phenols with aralkyl compds. Various examples are given, one of which involves the use of the product obtained by condensing 2-sulfo-4-chlorophenol with tetrachlorobenzyl chloride.

Mothproofing fur, wool, etc. I. G. FARBENIND A.-G. (Max Weiler, Berthold Wenk, and Karl Berres, inventors) Ger. 513,387, July 13, 1929, and 513,388, July 18, 1929. Addns. to 503,256 (C. A. 24, 5169). The materials are treated with asym. hydroxydi- or tri arylmethane derivs. obtainable by condensing *p*-chloro- or *p*-bromophenol (or their substitution products having one *o* position to the OH group unsubstituted) with aromatic hydroxy alics. or hydroxyhydrols or their anhydrides or derivs., which are themselves obtainable by condensing an aldehyde with one of the phenols mentioned. A suitable reagent is pentachlorodihydroxytriphenylmethanesulfonic acid, prep'd by condensing 2,4-dichlorophenol with the hydrol anhydride obtainable from *o*-sulfo-benzaldehyde and 2,4,6-trichlorophenol (513,387). Alternatively, the materials may be treated (1) with bis(hydroxybenzylated) aromatic hydroxy compds. having halogen in the *p*-position to the OH groups, but free from COOH or SO₃H groups, or (2) with nuclearly mono- or di-benzylated *p*-halophenols free from COOH or SO₃H groups and contg. no OH groups in the benzyl nucleus, or (3) with halogenation products of hydroxydi- or tri arylmethanes not contg. COOH or SO₃H groups (513,388). Cf. C. A. 25, 422.

Hydroxyarylmethanes. I. G. FARBENIND A.-G. Brit. 335,547, April 19, 1929. Mixts. of hydroxydi- and hydroxytri-arylmethanes, or mixed hydroxydi- and hydroxytri-arylmethanes, suitable for mothproofing wool, furs, etc., are prep'd by condensing, in an acid medium, one mol. proportion of formaldehyde with 2 mol. proportions of a mixt. of phenols comprising a *p*-halogenated phenol or phenols with other phenols, or by similar condensation of 1 mol. proportion of an aromatic aldehyde such as mixed *o*- and *p*-chlorobenzaldehydes or *o*-sulfo-benzaldehyde (but excluding aromatic hydroxy aldehydes and their sulfonic, carboxylic and sulfo-carboxylic derivs.) with 2 mol. proportions of a mixt. of phenols comprising *p*-alkylated phenols or their mixts. with other phenols, or a *p*-halogenated phenol mixed with other phenols (all the *p*-alkylated and *p*-halogenated phenols having free *o*-positions). Several examples are given and H_2SO_4 , AlCl_3 , FeCl_3 or a mixt. of H_2SO_4 and HIO_4 may be used as condensing agents. Cf. C. A. 24, 6036.

Rubberized cloth. E. I. DU PONT DE NEMOURS & Co. Fr. 693,890, April 14, 1930. Rubberized cloth is coated with a final layer of asphaltic varnish and heated to a temp. below the softening point of the asphalt. Cf. C. A. 24, 1524.

Apparatus for cleaning benzine used for degreasing and decoloring. MOISE DUVAL. Fr. 693,006, Mar. 29, 1930.

26—PAINTS, VARNISHES AND RESINS

A. H. SABIN

Studies in the painting of wood. I. Influence of wood structure on paint behavior. J. H. HASLAM AND S. WERTMAN *Ind Eng Chem* 23, 226-33(1931) —In order better to understand the relationships between wood and paint films, H and W examined microscopically the wood surface, the paint film and the interface of the 2. Differential staining of wood-paint sections made possible a study of the penetration of paint into soft woods of several species. The method is described in detail. Woods of apparent uniformity are revealed under the microscope as irregular. Differences in wood structure which have an important bearing on the behavior of the paint film include: (1) the sharpness of transition of spring wood to summer wood, (2) the distribution of pits on the fiber walls and their permeability, (3) the direction of the medullary rays with respect to the surface (flat- vs. edge-grained boards) and (4) the existence of resin ducts. A study of the shrinking and swelling of wood led to the discovery that the expansion of the summer wood is often greater than the over-all expansion of the wood. In such cases the swelling of summer wood is compensated for by compression of the spring wood. The expansion of the summer wood is less than the distensibility of a new paint film but exceeds many times the distensibility of a film 8 months old, hence the early loss of paint over summer wood. Penetration of paint liquids occurs through the pits and medullary rays. Penetration is uneven on flat-grained surfaces because of the existence of ray openings which occupy about 20% of the area. Uneven penetration results in nonuniformity of paint film and consequent early failure of paint film. Penetration of vehicles becomes more regular in the order: oil + turpentine, raw oil, bodied oil, varnish.

Effect of wood grain on paint durability. H. A. GARDNER, *Am Paint & Varnish Mfrs' Assoc. Circ* No 377, 123-30(1931) —A Pb-Zn house paint was applied over various primers to Ark. soft pine characterized by flat, angle and vertical grain. The primers were (I) the paint itself thinned according to directions, (II) Al spar varnish, (III) PbCrO₃ in oil and (IV) a graphite paint. Superior results were obtained on vertical grain with (IV).

Why some wood surfaces hold paint longer than others. F. L. BROWNE, U. S. Dept. Agr., *Leaflet* 62, 4 pp (1930) —It is not necessary for painters to change the composition of the priming coat for the different softwoods. Moisture gaining access to the backs of painted boards causes failure by blistering, peeling and discoloration. The painting characteristics of a board depend on the amount and distribution of the summer wood. The density of a softwood board measures roughly its ability to hold paint coatings. Cedars, cypress and redwood hold paint best and northern, western and sugar pines almost as well, whereas with southern pine and Douglas fir and western larch, serious flaking from the summer wood occurs sooner than with the former species. Other conifers lie between. Wood of the same species varies greatly in texture and density and therefore in painting characteristics. Knots and pitch give trouble. Edge-grain boards hold paint better than flat-grain. Southern exposure is the hardest on paint. Repainting should be done before the old coating has flaked or the bare wood is exposed.

Why wood painting research becomes a problem in forestry. F. L. BROWNE *J. Forestry* 28, 1136-45(1930), cf *C. A.* 24, 2313, 3913 —The variability in paint service on different woods is considerable. Density, ring width and type of grain, all widely variable, have a great influence on the integrity of paint films. Progress in wood-painting technology therefore concerns the forester. Results of tests described in previous papers are given.

Commercial paints, substitutes for white lead and anti-corrosion paints. A. VILA, *Chimie & Industrie* 24, 1052-67(1930); cf *C. A.* 22, 3998 —The results obtained after 4 yrs' exposure to the atm. are given, including photographs of the painted surfaces and of sections through them. Ti oxide and lithopone paints had the highest covering power, followed in order by ZnO, one of the white leads, mixed white pigments, Sb oxide and finally a white-lead paint. As regards resistance to weathering, the pure basic pigments (white lead, ZnO) which can combine with the acid radicals of the oil showed least loss to rubbing (light sponging with city or distd. water), while the dust thus removed from these paints had the highest content of org. matter; the Sb paint (Timonox) was already too badly deteriorated to make it worth while making these tests. Various accelerated aging tests for anti-rust paints were tried out, but apparently did not deteriorate the paint any more rapidly (in some cases less) than ex-

posure to the atm. or to fresh or sea water. As the value of these paints depends essentially on their impermeability, the following test is suggested. Apply 2 coats of the sample to be tested over the whole of a small Al plate and allow to dry normally. Fill the hollow of the plate with a soln. of NaCl and HgCl_2 in distd. water and let stand at atm. temp. if the coating is impermeable the soln. evap. without producing any corrosion, but if it is not impermeable the HgCl_2 reacts vigorously with the Al as soon as the soln. has penetrated to the metal. The tests showed that at the present time a no. of com. French anti-corrosion paints contg. no Pb salts are fully equal to, or better than the old minium paints.

A. PAPINEAU COUTURE

Technical analysis of titanium white. J. PFANHAUSER AND S. TOMASZEWSKI. *Przemysl chem.* 14, 333 5(1930).—The soln. contg. Ti is acidified with HCl with exclusion of air and then reduced with Zn for about 2 hrs. at 70–80°. High temp. causes the pptn. of insol. H_2TiO_3 . The soln. is then either titrated with 0.1 N KMnO_4 at 60–70° under CO_2 with exclusion of air or is reduced with an excess of ferrous NII, sulfate, which is then titrated back without the need of excluding air. A. C. Z.

Thickening of red lead. W. VAN WELLEN SCHOLTEN. *Farben-Ztg.* 36, 644–5 (1931).—The consistencies of pastes prepd. with (I) highly dispersed and (II) normal red lead and varying amts. of linseed oil were detd. at intervals during several months and plotted. The curves for I passed through a min. the breadth of which varied directly with the amt. of oil, 20% being the least amt. of oil for which the breadth was very great. The consistencies of II increased continuously from the time of their prepn. The settled layers of I were softer than those of II. In order to det. which oxide, PbO or PbO_2 , is responsible for the thickening properties of red lead, mixts. contg. varying amts. of these oxides were ground with linseed oil. The paste contg. 100% PbO_2 behaved like highly dispersed red lead and the mixts. with the highest percentages of PbO_2 had the softest settlings. In order to study the effect of the manner of prepn. the pastes, 84 parts red lead was ground with 9, 12 and 16 parts oil resp. After different intervals, the oil content of all the pastes was brought up to 16%. The results showed that softer pastes are obtained if only a portion of the oil be used at first.

G. G. SWARD

New methods of preparing chrome oxide green and hydrated chrome oxide green. A. C. HEINEMANN. *Farbe u. Lack* 1931, 6.—A few of the recent German patents for prepn. the chrome oxide greens are briefly reviewed.

G. G. SWARD

Survey of preparation of alkali chromates. ALFRED HEINEMANN. *Farbe u. Lack*, 1931, 20.—The author briefly reviews a few German and English patents on extn. of Cr from its ores.

G. G. SWARD

The toxic action of varnish, paints and polishes. DELFIN L. BARRIOS AND JOSÉ S. DEVOTO. *Semana méd.* (Buenos Aires) 1931, 1, 252–7.—A review is given of poisoning by vapors of solvents (PhH , PhMe , AcOEt and AcOC_2H_5).

A. I. MEYER

Spirit of turpentine in varnishes and paints. A. GÉRARD VAUDIN. *Bull. inst. pin* 1931, 4.—The merits of spirit of turpentine over white spirit are briefly discussed.

A. PAPINEAU COUTURE

Oiticica oil. G. G. SWARD. *Am. Paint & Varnish Mfrs' Assoc. Circ.* No. 377, 120–5(1931).—A white acetone-insol. material, which did not appear to be β -eleostearin, was obtained when oiticica oil was exposed in a thin layer to ultra-violet light. The yield was about 25%. A 25-gal. ester-gum varnish prepd. with oiticica oil compared favorably with one prepd. with tung oil. The literature of oiticica oil is briefly reviewed and a bibliography is given.

G. G. SWARD

Scheibler oil. G. G. SWARD. *Am. Paint & Varnish Mfrs' Assoc. Circ.* No. 377, 126–7(1931).—Some of the phys. properties of a drying oil prepd. from castor oil were: d_4^{25} 0.937, η no. (Wijs) 120.5, n_D^{25} 1.457, viscosity (G-ft scale) G, and color (Hellige) 11–1. When subjected to the Brown heat test, the oil did not gel in 2 hrs. The properties of a blown oil and of a stand oil prepd. from the raw oil are also given.


G. G. SWARD

Oxidation products of drying oils. I. β -Eleostearin from tung oil. R. S. MORRELL AND S. MARKS. *J. Soc. Chem. Ind.* 50, 27–36T(1931); cf. C. A. 24, 2901.—In the oxidation of the β -eleostearic glyceride the first product is a diperoxide, followed by a change to a monoperoxide, $(\text{CH}_2(\text{CH}_2)_2\text{CH}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}(\text{OH})_2\text{C}(\text{OH})(\text{CH}_2)_2\text{COO})_2\text{C}_2\text{H}_5$.



On methylation, Me esters, sol. and insol. in light petroleum, are formed. The max. number of methoxyl groupings which can be introduced is two. Attempts to methylate Me ricinolate were unsuccessful and resulted in the production of a methyl

ester of an acid isomeric with linoleic acid. The Me esters are oxidized by KMnO_4 to valeric acid and an intermediate C_{12} or C_{14} acid. Reduction by H_2 with Pd causes addition of two or four atoms of H. Treatment of the Me esters with HI or HBr causes reduction of the peroxide group and withdrawal of the methoxyl groups to form $\text{CH}_3(\text{CH}_2)_7\text{CH}(\text{CH}_3)\text{CHCH}(\text{CH}_2)_7\text{CO}_2\text{H}$. The evidence in favor of pos and neg



 polarity of the peroxide groups has been strengthened. The yellowing of drying films is due to rancidity, caused by disruption of the remote peroxide and the action of strong bases on the near peroxide to form colored salts. The inert character of the central double linking is very marked. There is a strong associating power of both peroxide groups. The peroxides are very stable in acid media. Polymerization of the oxidized esters is accompanied by a rise in refractive index. The loss in wt of linoxyn films is connected with degradation beginning at the remote peroxide grouping. A scheme of orientation on the lines of Staudinger's representation of polymerides is suggested.

H. M. STARK

Notes on aircraft finishes. H. A. GARDNER. *Proc Am Soc Testing Materials* 1930, Preprint No. 33, 37-45.—A discussion of various types of wing dopes. Acetate dope for priming coat evidenced lower tautness with good strength, while nitrate dope showed good tautness and satisfactory strength. Nitrate dopes properly applied and protected from light by Al-pigmented outer coats showed that the underlying coats of nitrate dope were not broken up by sunlight, the strength of the fabric was maintained and excellent tautness was effected. Two thin coats of clear spar varnish were applied over the Al pigmented dopes to make it smooth. Properly compounded nitrocellulose lacquer coatings are now being applied to the top side wing surfaces. The newly developed lacquer-varnish coatings usually contg nitrocellulose properly compounded in soln with phthalic anhydride glycerol resins appear to offer the advantages of quick drying, highly water resistant surfaces and exceptional durability. W. H. B.

The preparation of low-viscosity cellulosic colloids. CH. STARK. *Rev gén mat. plastiques* 6, 649-59(1930).—A brief review is given of the development of these colloids to meet the modern requirements of nitrocellulose lacquering and varnishing.

A. PAPINEAU COUTURE

Nitrocellulose lacquers. HANS WOLFF AND B. ROSEN. *Farden-Zig* 36, 730-7 (1931).—Ordinary kauri copal in either the natural or "run" states, tends to sep as a gel in films of nitrocellulose lacquers contg it. Solvent-extd kauri, however, is not subject to this defect. In hardness, abrasion resistance, flexibility, tensile strength and elongation, films of nitrocellulose lacquers contg extd kauri compared favorably with those contg ester gum. Films contg Manila copal were inferior to those contg shellac and extd kauri.

G. G. SWARD

Compatibility of resins with nitrocellulose solutions. HARRY C. NORMAN. *Ind Eng Chem* 23, 127-30(1931).—Although nitrocellulose, ester gum and AcOBu or AcOAm are compatible in all proportions and give clear films, many other systems give cloudy films. Of 132 combinations prepd by mixing nitrocellulose with 11 resins and dissolving the mixt in 1 of 12 solvents, 43 gave clear solns and films. 15 required only the addn of a plasticizer. 12 cleared up on the addn of a plasticizer and diluent, and 3 cleared up on the addn also of ester gum. 59 gave cloudy films in all tests. Some of the resins, notably vinyl acetate, rosin, dammar and Rezyl 12, are compatible with nitrocellulose in most of the solvents. Ester gum and Amberol H 9 could not be made compatible in systems contg diacetone alc., Et lactate, Cellosolve or Cellosolve acetate. Kauri and pontianac could be made compatible only in systems contg AcOEt, Cellosolve or Cellosolve acetate. Manila and shellac are incompatible in most systems and cumar could not be made compatible in any system. Incompatibility of resins with nitrocellulose in certain solvents is ascribed to lack of miscibility of resin and nitrocellulose rather than to ppts of resin by solvent. Careful examn of incompatible systems reveals the existence of 2 homogeneous phases but no ppt. Solvents and diluents belong to either an alc. or a hydrocarbon type and resins to either an alc.-sol. or a hydrocarbon sol type. In general, compatibility is most readily obtained by using solvents and resins of corresponding types, e.g., an alc. type solvent and an alc.-sol resin. However, in those systems requiring the addn of diluent or ester gum, it is found that the type of diluent or resin which produces compatibility is of the type opposite to that of the diluent or resin originally present. The ability of a plasticizer to clear up incompatible systems depends upon the nature of the plasticizer. A clear lacquer gives cloudy films when that type of solvent which makes for incompatibility is allowed to conc in the film during drying.

H. K. SALZBERG

Cativa resin. L. P. HART *Am Paint & Varnish Mfrs Assoc. Circ* No. 377, 131-2(1931).—A balsam-type resin from the tree *Prioria copaifera*, Griseb, made into a tung-oil varnish gave a product whose films were somewhat softer than those of varnishes made from the common resins. G. G. SWARD

The importance of glycerol and glycol in industry. G. PETROV. *Masloboino Zhiroroe Delo* 1929, No. 7, 15-21.—The new developments in the application of glycerol and glycol in the prep'n of synthetic resins by condensation with polybasic acids are reviewed, and attempts to prep' lannyn films' from the ethyl esters of fatty acids of linseed oil are described. The results of these expts were negative and P concludes that glycerol cannot be replaced in the drying oils by lower alcs. E. BIELOUSS

New method of determining the acid number of copals. ZYGMUNT LEPPER. *Przemysl Chem* 15, 1-5(1931).—One g finely powd copal is boiled in 20 g terpeneol under a reflux until dissolved. Some difficultly sol copals may take 1-3 hrs. The hot soln is diluted with 20 cc $C_{12}H_{22}O_{11}$ and on cooling is titrated with 0.1 N KOH. With difficultly sol copals the hot terpeneol soln is treated with 20 cc $C_{12}H_{22}O_{11}$ and 20 cc alc 0.5 N KOH, and heated for about 15 min. Since the soln turns dark it is diluted with 200 cc $C_{12}H_{22}O_{11}$ and after cooling titrated back with 0.5 N H_2SO_4 against phenolphthalein. Compared with the methods of Gardner, Dietrich, Marcusson and Winterfeld this method has the advantages in that it gives consistent results in a series of detns on the same copal, it is general and does not require a different set of solvents for every kind of copal, the results obtained are very probable. The acidity values obtained by this method on various kinds of copals may be used for comparing their acidities. Tables of the acidity values of many kinds of copals are given. A. C. Z

Production of oleoresin from *Pinus halepensis* of Spain. VICENTE CUTANDA AND MARIANO SEVILLA. *Servicio Forestal de Investigaciones y Experiencias* 2, No. 5, 29-42(1929).—The production of oleoresin from stands of *P. halepensis*, which are found along the Mediterranean coast, began during the World War. At present about 600,000 trees of this species are being turpentine-d. The oleoresin is being distd at 7 different plants. The yields per tree per yr vary from 1.05 to 4.50 kg, depending upon growing conditions. These yields compare favorably with those from the more commonly worked *P. pinaster* (maritime pine). The compn. of the oleoresin as given by Lacue (C. A. 23, 5037) is discussed. The resin is somewhat low in grade because of a greenish tinge. Cleaning the gum before distn, according to the usual Spanish practice, is especially advantageous on account of the presence of much tannin. Water is sepd. by decantation from the oleoresin before distg by allowing the mixt to stand in closed vessels for 10-12 hrs. The resin is invariably sunned to improve its color. The turpentine must be distd to remove its greenish color. A proposed alternative is filtration through $(CO_2)H_2$. H. K. SALZBERG

Azo dyes [for red varnishes] (U. S. pat. 1,791,444) 25.

Paint composition. RICHARD L. BRADDOCK and DALE F. McFARLAND. Can. 308,612, Feb. 17, 1931. Turpentine is added to paint and then there is added a soln made up of water, shellac and $NaHCO_3$.

Weather-resisting paints. I. G. FARBENING A-G. Brit. 335,626, May 2, 1929. There is incorporated with paints or other coating materials at least such a quantity of pptd. ZnO (dried by heating to a temp. not materially exceeding 500°) as is necessary to combine into soaps all free or combined fatty acids present in the binding agent such as linseed oil or China wood oil. Various details and examples are given.

Roof paint. JOSHUA D. WALSTON (to Black Diamond Paint Co.) U. S. 1,791,455, Feb. 3. A paint is formed from ingredients including coal tar, rosin, PbO , $ZnSO_4$, $CaSO_4$, gasoline and kerosene.

Paints, stains, etc., comprising colloidal dispersions. SAMUEL CABOT (to Samuel Cabot, Inc.) U. S. 1,791,119, Feb. 3. Pigment particles of colloidal fineness are carried in suspension in an org. dispersion medium such as linseed oil and turpentine or solvent naphtha together with a semi-plastic treated oil such as a linseed oil jelly.

Paint filter. RUDOLF SCHMID. Ger. 515,439, Sept. 28, 1929.

Indelible ink. MARCEL ROURE. Fr. 692,929, June 28, 1929. Indelible ink is made by dissolving salicylic acid 2, gum arabic 25, alum 100 and oxalic acid 50 g. in 1 l. of water, boiling and adding 100 cc of ox gall and 100 cc of alc to make soln. A. Aniline dyes according to the color of ink desired are added to a boiling soln. of 20 cc. of HCl in 1 l. of water to form soln. B. Solns. A and B are mixed, allowed to stand and then filtered.

The kernels yielded 61.7% oil (on the dry basis), which had const. agreeing well with those recorded by Bolton for stillingia oil. A PAPINEAU COUTURE

Some developments in the oil palm industry. H. M. LANGTON. *Ind. Chemist* 7, No. 73, 71-4 (1931) E. H.

The lipase of olives and olive oil. E. PANTANELLI AND S. VERDESCA. *Rend. accad. sci. Napoli* [3], 36, 76-83 (1930). —The acidity which develops in olives after they are picked is due to a lipolytic enzyme which accelerates the hydrolysis of the glycerides into free fatty acids and glycerol with absorption of water. The activity of the lipase increases rapidly with the destruction of the cellular structure. The greater the initial acidity of the olives at the time of pressing, the more apt is the oil to be rich in lipase and develop more acidity. The acidity which develops in the oil is proportional to the water which remains emulsified during the process and also proportional to the glycerol already present. The glycerol, therefore, is the factor which governs the course and intensity of acid development, since it introduces water and lipase into the oil. PETER MASUCCI

Influence of bagging on the percentage of oil in castor beans. G. B. PATWARDHAN. *Poona Agr. Coll. Mag.* 22, 13 (1930). —Bagging castor bean flowers for seed propagation had no significant effect on the oil content of the seed. K. D. JACOB

The determination of the oil content of seeds in series. A. LEBEDYANTZEV AND N. DMITRIYEV. *Maslobosno Zhivoe Delo* 1929, No. 10, 3-10, No. 11, 3-9, No. 12, 9-17. —With the object of devising a method for serial detn. of oil in seeds L and D investigated the influence of various factors on the extn. of oils. After 4 hrs of extn. 34.65% of the oil was obtained, after 18 hrs., 41.65%, and after 36 hrs., 41.76%. The usually specified time, 5-6 hrs. for extn. with a Soxhlet, is much too short, about 18 hrs. would be required for practical purposes. With 30 siphonings per hr. the extn. was complete in 16 hrs., as compared with 36 hrs. for 6 siphonings per hr. By pulverizing the material to a very fine state and with 30 siphonings per hr. the duration of the extn. can be reduced to 6 hrs. By previous maceration of the material with Et_2O for 18 hrs. the extn. time was reduced to 2 hrs. Up to 2 g. of a finely powd. material can be completely extd. within 6 hrs. at 30 siphonings per hr. Three g. of the material under the same conditions gave results which were lower by 1.1%. The Et_2O must be freed from alc. and dried over CaCl_2 . To avoid the contamination of the Et_2O with H_2O the material should be dried at 100° . Drying oils should be dried in a N or a CO_2 atm. to avoid O absorption. Instead of drying, the material may be mixed with calcined CaSO_4 . The best method for drying the extd. oil is to heat it in a N or a CO_2 atm. not over 60° . The oil content was detd. from the wts. of the extd. residues. The app., representing essentially an enlarged Soxhlet, was charged with 20 thimbles, each contg. 2 g. of material. The thimbles were allowed to soak for 24 hrs. in Et_2O and the extn. was carried out in 3 hrs. at 6 siphonings per hr. Uniform results were obtained agreeing within 0.16% with individual extns. in Teuge's app. The subsequent drying of the residues presents no difficulties, as it can be carried out by heating in the air at 100° . E. BIELOUSS

The relationship between acidity and titer of oils. I. STETENKO AND I. PANTELEVYEV. *Maslobosno Zhivoe Delo* 1929, No. 10, 20-1. —Although the titer (i. e., solidification point) of the fatty acids is higher than the titer of their glycerides, a high titer is no indication of a high acidity, as the neutral oil forms a eutectic mixt. with the free fatty acids. Thus a sunflower oil with an original acidity of 3.5% had a titer of 35.3° and 46.1° for the free acids. In admixts. up to 30% free acids the titer was below 35.3° . A hydrogenated sunflower oil with a titer of 56.4° and 61.6° for the free acids had a lower titer in admixts. up to 60% free acids. E. BIELOUSS

The splitting of oils and fats nearly to completion. G. PETROV AND A. PICHUGINA. *Maslobosno Zhivoe Delo* 1929, No. 9, 53-6. —By the use of the naphthenic sulfonic acids oils and fats were hydrolyzed up to 97.87-99.6%. The hydrolysis was carried out in two steps by adding the sulfonic acids (0.9 to 1% of the oil) in two portions. E. BIELOUSS

Denaturing of proteins of soy bean by alcoholic extraction. L. KOJI, OKANO AND MAYORU NIJOMIYA. *Abstracts from Rept. Central Lab. S. Manchuria Railway Co.* 1929, 7-9. —Alc. extn. decreases the H_2O -sol. and NaCl soln.-sol., and increases the NaOH -sol. protein. No difference in the amino acids of the original and extd. products was found. V. F. HARRINGTON

Extraction of soy-bean oil with alcohol. II. 1. Solubility of soy-bean oil in alcohol. MASANORI SATO AND HIROSHI SAKAI. *Abstracts from Rept. Central Lab. S. Manchuria Railway Co.* 1929, 1-3. —The soly. of soy-bean oil in alc. of various concns.; in alc. dild. with C_6H_6 , benzene or oleic acid, in MeOH , and in MeOH dild. with C_6H_6 ,

has been detd. 2. Experiments on the extraction of soy-bean oil with alcohol. MASANORI SATO AND MITSUO YOKOCHI *Ibid* 3-6—The optimum conditions are 2 extrns with alc. (96% by vol) for 2 hrs. at 75° or above. The ext. cooled to 25° deposits the oil. Extn above the b. p. gives better results. An app. is described. Yields are equal to those with benzine, the product is edible. 3. Separation and recovery of phosphatides and carbohydrates from the alcoholic extract. MASANORI SATO AND IMAWOTO SETO *Ibid* 6-7—The residue from the alc. phase contains 27% carbohydrates, 1.27% N, and 1.07% P. Alc. dissolves the phosphatides. They can be coagulated by diln of the alc. with H₂O and addn of acid or salts. V F HARRINGTON

Determination of sandy matter in soy-bean cakes. SHUNICHI USAMI *Abstracts from Rept. Central Lab. S. Manchuria Railway Co.* 1929, 23-4—The cake is shaken with CCl₄ and the sand weighed or measured in a flask similar to an inverted Babcock flask. V F HARRINGTON

The soy-bean oil from Kuban. M. SEAGEEV *Masloboino Zhirovoe Delo* 1929, No. 3, 47-51—Fifty-five samples of the 1927 crop were analyzed. The n_D^{20} , and the acid and I nos. of the Kuban oil are somewhat lower than the corresponding consts. recorded for the oils of different origins. The yield of beans per hectare depends on the vegetation period; it reaches 2250 kg. for the late crops (140-163 days). The oil content varies from 19.6 to 25.3%. F. BIELOUSS

Examination of Japanese beeswax. I. General properties. II. IKUTA *Chem. Umschau Fette, Oele, Wachse u. Harze* 38, 7-10 (1931). d 0.956, n_D²⁰ 1.4557-1.4560, acid no. 7.5-8.4, ester no. 75.6-79.2, sapon. no. 83.1-84.6, I no. (Wij's) 11.3-14.0, acetyl no. 20.7-18.7, unsaponifiable constituents 58.3-59.5%. Only 2 of the other 6 samples were considered pure wax (European type of bees). d₄¹⁰ 0.8141-0.8132, m. p. 62.0-63.5°, n_D²⁰ 1.4554, acid no. 19.8-19.4, ester no. 74.0-74.1, sapon. no. 93.8-93.5, I no. (Wij's) 8.5-7.2; acetyl no. 13.1-14.4, unsaponifiable constituents 49.5-48.6%. In addn, I tabulates the consts. of the saponifiable constituents of the waxes, their ales and their hydrocarbons. P. ESCHER

Manganese soaps. HANS WAGNER AND G. HOFFMANN *Farben-Ztg.* 36, 689-93 (1931)—The reaction of umber with linseed oil fatty acids was studied by exposing mixts. of umber and fatty acids to ultra-violet light, daylight and darkness. The samples were exposed on microscope slides with and without cover glasses. In all cases where crystn. took place, ether-sol. Mn compds. were found, but the formation of Mn soaps was not definitely proved. Pronounced crystn. took place only in the presence of oxidizing or catalytic compds. Thus pyrolusite produced no crystals, but several ter- and quadrivalent Mn compds. did so as a result of the O₂ liberated in the reduction of the Mn to the bivalent form. G. G. SWARD

Analysis and preparation of medicated soaps. G. MATOLCSA *Magyar Gyógyszerésztud. Társaság Ertetöje* 5, 489-502 (1929)—The flame reaction of soaps may show what alkalies have been used in their prepn. Color, odor and consistency should be observed. The water content can be measured on the basis of the resistance against cutting. The paste-cutting machine of Rejtő has been used for this detn. Fatty acids, content of alkalies, and org. and inorg. filling materials are detd. according to standard methods. A short description of the prepn. of medicated soaps is given. S. S. DE FINÁLY

Combinations of raw materials for toilet soap. BUNJI NAKAGAWA *J. Soc. Chem. Ind., Japan* 34, Suppl. binding 22-7, 28-31 (1931)—Soaps of 8 different oils were mixed in proportions of 10-80% with a soap consisting of 80% Na tallow soap and 20% Na coconut-oil soap and the lathering no. (L. N.), lathering vol. (L. V.), lathering coeff. (L. coeff.) and rate of lather extinguishment detd. L. coeff. = L. N. × L. V., and the rate of lather extinguishment = {(L. coeff. at 1 min. after shaking) - (L. coeff. at 5 min. after shaking)} / (L. coeff. at 1 min. after shaking) × 100. Na palm-oil soap (I) has a good effect upon lathering. Hardened soy-bean oil soap (II) of 67.7 I no. has a better effect at 40° than at 20°, where it shows a reverse tendency. A Na lard soap (III) has a good effect on the L. coeff. at the start, but the rate of lather extinguishment rises with the increase in amt. of this soap. A 40% mixt. is the best for lathering power. Na hardened fish-oil soap (IV) decreases the L. coeff. The Na soap (V) of Chinese

vegetable tallow has not much effect on the L coeff., but the rate of lather extinguishment is larger. At 20° the L coeff. decreases. Na castor-oil soap (VI), Na peanut-oil soap (VII) and K tallow soap (VIII) all have a good effect upon the lathering, but VI is the most desirable. The soly of mixed soap was measured in an attrition app composed of an endless rubber band rotating at 1080 cm per min and dipping into a water bath at 40° and 20° . A soap cake $2 \times 3 \times 2$ cm was fastened and weighted with 135 g so that the rubber band could rub it. The duration of the operation was 10 min at 40° and 20 min at 20° , and by weighing back the soap the attrition was detd. The degree of attrition is designated "rubbing soly" and is calcd by the formula: rubbing soly = $100 \times \{ \text{soap by attrition (g)} + \text{rubbing surface (cm}^2) \}$. The rubbing soly of Na tallow soap mixed with 20% Na coconut oil soap is smaller than that of the former alone but increases with increase of the latter. The mixt. with I or II does not increase the rubbing soly. In the mixt. of III the rubbing soly increases with increase in amt. of this soap. In mixts. of IV it decreases with increase of this soap. In the case of V the rubbing soly tends to decrease, and this effect is greater at 20° than at 40° . A 2.5% mixt. of VI increases the rubbing soly, as do also VII and VIII, 15% of VII being equal to 2.5% of VI. The relative hardness of the mixed soaps was detd from the distances 2.5-cm metal points 2 and 3 mm in diam (weighted to a total of 300 and 600 g, resp) penetrated into the soaps, a homogeneous curd soap being the standard of comparison. The hardness of mixed Na tallow and Na coconut-oil soap increases with the increase in the amt. of the latter until, with 40% coconut-oil soap, the soap is unsuitable for toilet use. I and II (I no 67) have a tendency to soften the mixed soap, I having a greater effect than II. IV (I no 22.3) in lower percentages produces a slight softening in the mixt., but when 50% is used, it hardens the product. V increases the hardness as its amt. is increased. VII, VIII and VI (effective in this order) soften the mixt. according to the amts added. Conclusion: palm oil, hardened soy bean oil and Chinese vegetable tallow can generally be used as substitutes for tallow, and the most desirable for the combination are castor oil, peanut oil and the K soap of tallow.

F. SCHERITTEL

The utilization of soap lyes. S. LIBERMAN and A. KORNBLIK. *Mashobino Zhiroroe Delo* 1929, No. 6, 21-5.—The soap lyes left after the salting out of the soap base, contain about 1.3% NaOH and 0.5% Na_2CO_3 . L. and K. recommend their recovery by treatment with free fatty acids and rosin.

E. BIELOTSE

Actual problems in the manufacture of washing compounds. J. LEINBÖRGER. *Seifensieder Ztg* 57, 869-72 (1930).—L. discusses the economic and tech. phases of toilet, laundry and soft soaps. The brown spots on cottonseed-oil soaps may be prevented by hydrogenating the oil to hard consistency or by polymerizing it at $200-280^\circ$ for 5-6 hrs in a CO_2 atm. before sapon.

P. FETTER

Ambergris and how to recognize it. F. R. MORRISON. *Bull. Tech. Museum Sydney* No. 15, 9 pp (1929).—The different varieties of ambergris are described, together with their values, uses and characteristics. All ambergris will float on sea water, and if the latter is slowly heated the material will melt to a brown or black liquid before the water reaches its b.p. A piece of ambergris about the size of a pea is almost completely sol. in about $\frac{1}{2}$ oz. of warm alc., giving a fluorescent soln. A needle, previously heated in a candle flame for 10-20 sec., on being pressed into a sample of ambergris to a depth of $\frac{1}{8}$ in. forms a dark brown resinous liquid which when touched by the finger leaves pitch like "strings" adhering to the skin. If the needle be withdrawn and placed in the flame, the ambergris burns, if the flame be extinguished the odor of the "smoke" is somewhat fatty or resinous and resembles that of burning rubber.

B. C. A.

Oiticica oil (SWARD) 26. Insect oils and fats (TIMON-DAVID) 111. Value of ash determinations of fatty oils (SCHARFER) 22.

Apparatus for washing and mixing fats and waxes having the consistency of butter. RICHARD BENDLIN and OTTO KAR URBASCH. *Pr.* 692,687, Jan. 6, 1930.

Coloring higher fatty acids such as stearic acid. HERMANN SCHLADENBACH and HERBERT HÄHLK (to General Aniline Works). U. S. 1,791,431, Feb. 3. A coloring substance such as a thioindigo and an amino compd. such as triethanolamine are added to fatty acids while the latter are in liquid state. Numerous examples of use of Cu salts, etc., also are given.

Extracting oils and fats. STRAUSS & CO. (Ernst Müller, inventor). Ger. 515,058, July 20, 1926. This relates to the extrn. of oils and fats from oil seed press cakes, used bleaching earth and like residues. The material is ground and mixed with

a solid diluent e. g., sawdust and the mixt is lightly pressed into briquets which are then extd. with a solvent in known manner

Apparatus for extracting oils from seeds, sewage or other materials by solvents. CLARENCE J BLEIL U S 1,791,398, Feb 3 Various details are described of an app. comprising an endless filter made up of connected sections Cf C A 24, 1257

Recovering dry palm oil from its admixtures with water. JAMES H HOLDEN and JOSHUA C WHEZEL (to American Sheet and Tin Plate Co) U S 1,790,748 Feb 3 A water in palm oil mixt sep'd from a palm oil in water emulsion formed by washing palm oil from tin plate after it passes from the tinning pot, is heated until substantially all the water is vaporized, the mixt of oil and vapor resulting from the heating is then delivered into a draft of air of sufficient velocity to carry away the vapor and of insufficient velocity to pick up the oil App is described Cf C A 24, 3749

Soap PIERRE P A CHAPPELLE Fr 693,580, April 9, 1930 A soap for domestic use or for mechanics contains tallow soap 50, fine white pumice 43 glycerol 5 and resin 2%

Soaps, creams, etc. Soc d'Études Scientifiques et Industrielles Fr 692,454 June 19 1929 Biol products are used in soaps, creams, etc Fr 692,455 describes the use of exts from fungi etc in soaps, creams, etc

28—SUGAR, STARCH AND GUMS

J K DALE

Electrical conductivity and ash content of sugar house products. K ŠANDERA *Bull assoc chim sucz dist* 47, 446-9(1930)—A new app is described for the electro-metric detn of ash in which Pt black electrodes are eliminated An optical method is used to det. the end point instead of a galvanometer The app is in principle a Wheatstone bridge Two of the 4 resistances are lamps, the light of the lamps falling on the surface of a triangular prism The bridge is balanced when the 2 lamps appear to be of equal brightness The prism reflects the light from both lamps onto the field of vision, thereby facilitating comparison The third resistance is a const., and the fourth is a variable resistance consisting of the sugar soln with 2 electrodes The distance between the electrodes is variable and is changed to balance the bridge The distance apart may be calibrated to read ash content F. CAMPS CAMPINS

The action of nonsugars in sugars upon the caramelization test. J PUCHERNA. *Listy Cukrovar* 49, 13-20(1930).—Thin-walled containers contg 6.5 g of a sugar mixt. were dried at 100° for 2 hrs. The containers were then submerged in an oil bath at 170° for 15 min., withdrawn and cooled in the air The residue was dissolved in 50 cc. water, filtered and compared in an objective photometer The remaining soln was treated with 0.5 cc. 2 N NaOH, digested over a water bath for 10 min., cooled and compared in the photometer for additional coloration. The action of salts falls into 3 classes (1) large primary and small secondary coloration, (2) small primary and large secondary coloration, (3) very little coloration Salts falling into class (3) are NaCl, KCl, BaCl₂, K₂SO₄, Na₂CO₃, AcONa, AcONH₄, Na oxalate All salts in this class except AcONH₄ are strongly electropositive and their solns are neutral or alk The small coloration in alkali formed by the chlorides and acetates indicates a decompn of sucrose into glucose and fructose which readily give rise to colorations Acetates form more primary coloration than chlorides Sucrose in soln and at high temps behaves as an acid At 170° AcOH is volatilized from AcONa by sucrose The sucrose liberates an equiv quantity of an acid from the salt of the acid The action is most pronounced for weak acids, the resulting soln is alk., which favors a decompn of sucrose For NaCl, KCl and BaCl₂, the liberated HCl is small, the medium is neutral and no decompn due to increased alk. occurs AcONH₄, as the salt of a weak acid and base, remains neutral in action Na₂CO₃ leaves an alk. medium, and some coloration results K₂SO₄ decomposes at 1070°, it cannot be attacked by the sucrose and remains neutral during the decompn Na oxalate behaves more like NaCl than AcONa (COOH) is a stronger acid than AcOH and the decrease in coloration agrees with their dissocn. consts Class (1) consists of salts or amides of amino acids Metallic salts of amino acids have a greater effect than the corresponding amides because of the action of the alkali Class (2) forms products which become colored in alkali NH₄Cl, Fe(NH₄)₂(SO₄)₂, NH₄ oxalate and (NH₄)₂SO₄ cause a large primary coloration in addn to the secondary; CaCl₂, betaine-HCl and glutamic acid-HCl give very little primary coloration, Fe₂(SO₄)₃ and KHSO₄ lie between these groupings.

Betaine-HCl decomposes at 228°; glutamic acid HCl, at 208°, in the presence of sucrose the decompn. begins at a lower temp. The effect of sucrose upon salt is marked. HCl is liberated and attacks the sucrose, forming invert sugar. The additional coloration in an alk. medium is a linear function of the catalytic salt. Too much acid, 0.2%, causes a more thorough decompn. with the production of ulmic acid and caramel. CaCl_2 , although similar to BaCl_2 , acts as a strong dehydrating agent and falls into class (2). The larger primary coloration with Na_2SO_4 and $\text{Fe}_2(\text{SO}_4)_3$ is due to the dehydrating and oxidizing action of H_2SO_4 . The large primary coloration with NH_4Cl , $(\text{NH}_4)_2\text{SO}_4$, NH_4 oxalate and $\text{FeSO}_4 \cdot (\text{NH}_4)_2\text{SO}_4$ is due to the action of NH_3 on sucrose, with NH_4Cl the action of NH_3 exceeds that of HCl. The class (2) catalysts form a fluid mixt. very soon after heating. Larger addns. cause foaming with intense odors suggesting $\text{C}_6\text{H}_5\text{CHO}$. The solns. become acid. Sucrose may act as a solvent for salts capable of dissoen. Granules of salts are often surrounded by a low melting mixt. which contains sucrose in a fluid state. The sucrose may exert its high dielec. properties as a solvent, the dissoen. of salts in this "sucrose solvent" is similar to dissoen. in water. A mixt. of salts in the form of molasses was also tried. The effect is complicated not only by the presence of mixed salts but also by the colored org. substances of molasses whose effects have not been studied. The primary coloration with molasses is large, the secondary is small; the effect is due primarily to substances of class (3). FRANK MARESH

Iodometric determination of sugar. M. S. FILOSOPOV. *Nauk. Zapiski Tsukrov.* *Prism.* 9, 586-9 (1930).—In certain cases the sucrose cannot be detd. polarimetrically. The following method may be used on molasses. The normal wt. of molasses is dissolved in water to 130 cc. Fifty cc. of soln. (10 g. of molasses) is transferred into a 100-cc. flask and 25 cc. of H_2O contg. 5 cc. of HCl (d 1.19) is added. The inversion is conducted according to the Herzfeld method. The soln. is made up to the 100-cc. mark and thoroughly mixed. Twenty cc. of the inverted soln. is transferred to a 400-cc. flask, 20 cc. of Na_2CO_3 soln. (84 g. per l.), 1-2 cc. of lead acetate and the same amt. of $\text{Al}(\text{OH})_3$ are added. The excess of lead acetate is eliminated with soda ash and the whole made up to the 400-cc. mark and filtered. Fifty cc. of filtrate is transferred into a 250-cc. Erlenmeyer flask. Four to five g. of pure sucrose, and 50 cc. of Fehling soln. are added. The mixt. is boiled for 2 min. and then cooled by addn. of 100 cc. of H_2O , transferred to a 250-cc. flask and made up to the mark. One hundred cc. is transferred to an Erlenmeyer flask, 5 cc. of a soln. contg. 0.65 g. $\text{KCNS} + 0.1$ g. KI in 5 cc. of water, is added and thoroughly mixed. Ten cc. of 6.5 N H_2SO_4 is added, mixed and titrated with 0.1387 N $\text{Na}_2\text{S}_2\text{O}_5$, $5\text{H}_2\text{O}$. Five cc. of 1% starch soln. is added as indicator. The Bunsen table gives mg. of invert sugar and this multiplied by 0.9 gives the sucrose. Cf. C. A. 24, 985. V. E. BAIXOW

A preliminary study of the glucose, sucrose and refractometer solids relationships of five sugar cane varieties grown under Laguna conditions. MANUEL L. HOLLERS. *Philippine Agr.* 19, 299-305 (1930).—The sucrose content varies widely in different parts of the plant. The refractometer solids vary directly with the gravity purity. In all varieties the refractometer solids and gravity purity decrease and the glucose content increases from the bottom to the top of the plant. A. L. MEHRING

The relation between the ash content of raw sugar and the conductivity of sugar solutions of different concentrations as well as of acidified sugar solutions. O. SPENGLER AND F. TÖDT. *Z. Ver. deut. Zuckerind.* 80, 553-65 (1930).—Expts. on some 60 samples of German raw sugars show that the elec. detn. of ash is more accurate at 5° Bx than at a higher Bx (e. g. 25° Bx). The correction factor of Zerban for the decrease of the difference between the chem. and the elec. ash detn. was applied to 70 samples of German raw sugars. The results were adequate. This shows that the non-sugars in the samples tested by Zerban are different from the non-sugars in the German samples. F. CAMPS-CAMPTNS

The Kämpf viscometer. ERHARD LANDT. *Z. Ver. deut. Zuckerind.* 80, 949-55 (1930).—A description is given and the method of operation of the app. is reviewed. Tests were carried out on various sugar solns. Conclusion: It is sufficiently accurate for solns. of viscosities over 3 poises. (There is no turbulent flow above this value.) It is especially recommended for solns. of high viscosity. The instrument will work even where solid suspended particles are present. The viscosity in this case is, of course, not the same as the viscosity of the pure liquid (cf. C. A. 24, 1767). F. C.-C.

The determination of reducing sugars in the presence of sucrose by potassium ferrieyanide. R. HARMY. *Bull. assoc. chim. sucr. dist.* 47, 385-7 (1930).—The method of Jonescu and Vargolici (cf. C. A. 15, 3564) is reviewed. Comparative tests were made on sugar solns. with varying contents of sucrose. It is shown that CuSO_4 -Rochelle salts methods give results which are too high, increasing with increasing amts. of sucrose.

For the J and V method the true result is obtained as long as the sucrose content is below about 30%. The end point is quite sharp. F CAMPS CAMPINS

The continuous recording of the alkalinity of saturation beet juices by conductivity measurements. O SPENGLER AND F TÖDT. *Z. Ver. dest. Zuckerind.* 81, 1-12 (1931).—The alk. tester of Reinhard Lardner which gives a continuous indication and record of the alk. by cond. measurement was tested in a factory. Such control seems feasible for continuous satn. A conclusive decision is possible only after further investigation, but a straight line relationship is indicated. The influence of juice composition and needs to be studied. F CAMPS CAMPINS

The Sándera objective photocolormeter. ALON DOLINIK. *Licht Colour* 49, 203-5 (1930).—Using various liquors available in a refinery, D. compared the Stammer and Sándera photocolormeters. The av. variation in readings with the Stammer was 3.5-5 times as great as with the Sándera colormeter. The variation of eyes is insignificant with the Sándera colormeter but large with the Stammer. F M

Effect of certain non-nitrogenous substances on the polarimetric determination of sugar. F PARIS. *Jed. saccharif. ind.* 23, 217-9 (1930). *Chimie et industrie* 24, 1932 (1930).—P. shows that the polarimetric detn. of sugar can be vitiated by soln. of foreign substances derived directly or otherwise from pectin. Heating dried beet pulp with distd. water on the water bath gives a strong d rotatory soln. because of the presence of arabinose (arabinose anhydride), which is partially pptd. by basic $Pb(OAc)_2$, if $Ca(OH)_2$ water is used instead of distd. water, the soln. is strongly l rotatory on account of the presence of Ca or Mg salts of pectic acid, if a soln. of basic $Pb(OAc)_2$ is used, the soln. is much less l rotatory. A PAPINEAU COURTNE

Determination of the amount of crystal in massecuites by the electrical conductivity method. K. SANDERA. *Z. Zuckerind. Centralblatt Rep.* 54, 311-4 (1930).—In place of calcn. based on formulas employing the dry substance or polarization of the massecuite, elec. cond. measurements afford a more reliable method for the calcn. of the amt. of crystal to be expected, and are besides considerably more expeditious. B C A

Analysis of sugar cane. E. HADDOX. *S. African Sugar J.* 14, 675 (1930).—Sucrose in cane is usually estd. by multiplying sucrose in juice by a factor which necessarily varies with the proportion of juice in cane. H. proposes to det. sucrose % cane directly by repeatedly extg. on a Buchner funnel 100 g. of finely disintegrated cane, neutralized by means of lime water, with hot water until a total vol. of 100 cc. is obtained. The ext. is clarified, polarized in a 400-mm. tube, and the result multiplied by 1.3. To obtain juice % cane, a quantity of the pulp is pressed in a tannery press; the extd. juice is clarified and polarized, and its % gr. is detd. Then juice % cane = 100 sucrose % cane / sucrose % juice. The quantity $(100 - \text{juice \% cane})$ is termed marc, which is the hydrated insol. part of the cane. The water % cane is detd. by drying the pulp. The fiber % cane is then found by subtracting from the marc the difference between water in cane and water in juice. Six typical analyses are given, showing that the fiber % cane detd. by luviation is considerably lower than that found by H.'s method. F W ZERRAN

Determination of fiber and sucrose in cane. M. REICHARD. *S. African Sugar J.* 14, 743 (1930), cf. preceding abstract.—Haddox's article is criticized. The inaccuracies inherent in cane sampling may lead to errors of 2 units in the % fiber. Even if a correct sample of pulp can be obtained the analysis should be made by pressing the pulp, analyzing the juice for Brix and sucrose, and the residue for water and sucrose. From these figures the sucrose and fiber % cane can be calcd. directly. H.'s methods are based on incorrect assumptions, and the results for fiber will be accurate only within -2.4 and $+5.4$ units. H.'s figures actually show errors of this magnitude, as the water held by the dry fiber varies in his results from 31.8 to 86.2%. The problem of fiber detn. in cane has not yet been solved satisfactorily. F W ZERRAN

A new spectral colormeter. O SPENGLER AND E. LANDT. *Z. Ver. dest. Zuckerind.* 81, 15-24 (1931).—A detailed description and method of use of the app. are given. In principle it is like a Duboscq colormeter equipped with a monochromator. Readings are made at $\lambda = 610, 560$ and 480 . Exptl. figures show the accuracy of the instrument when used as a Stammer-app. in the broader sense and also for the comparison of 2 colored solns. In the latter case $a_1/a_2 = c_1/c_2 \cdot h_1/h_2$, where a is the extinction coeff. at the given wave length, c the concn. and h the height of the liquid column. F CAMPS CAMPINS

Apparent disagreement in the analysis of molasses. J. J. WETES. *Licht Colour* 49, 215-4 (1930).—The molasses of the present season have been very dark and have

led to discrepancies in analyses. All discrepancies have been due to improper sampling of molasses for analysis. Studies in sampling are in progress. FRANK MARESH

Materials for the study of sucrose crystallization. I. A. KUKHARENKO AND I. N. KAGANOV. *Nauk. Zapiski Tsukrosos. Prom.* 9, 402-9 (1930).—The rate of crystal growth has the highest velocity during the seeding at the beginning of the process. The maximum speed of crystal growth does not exceed 10000 mg. on 1 sq. m. of crystal surface per min. One hr. from the moment of the seeding is required to obtain crystals of 1 mm. length. For a sugar crystal of an ordinary kind the approx. size is $P = 0.35 L^3$ and $S = 2 L^2$, where P is the wt. of the crystal in mg., S its surface in sq. mm., and L its length in mm. V. E. BAIKOV

Influence of bagasse analysis on the sucrose content of the cane. E. HADDON. *S. African Sugar J.* 14, 741 (1930).—Previous investigators have shown that prolonged boiling and greater diln. increase the polarization of the bagasse ext. This is probably due to the extn. of optically active non-sucrose substances derived from the hydrolysis of hemicelluloses. To prevent this, the extn. must be carried out in the presence of $Ba(OH)_2$, alkyl to phenolphthalein being kept up until the end of the digestion. The baryta method applied to bagasse gave results 0.08 to 0.3% lower than the ordinary method. F. W. ZERNAN

The deterioration of Philippine sugars under varying degrees of humidity. QUERINO D. RENDON. *Philippine Agr.* 19, 383-90 (1930).—Increase of H_2O in cane sugar favors rapid growth of organisms, which causes the sugar to deteriorate and introduces an error in the polarimetric detn. of sucrose. The max. degree of humidity in which sugar will not deteriorate is 60.2%. In an atm. of 100% relative humidity 30% of the sugar was lost in 4 months. A. L. MENZING

Adsorption from sugar solutions. III. Adsorption in the carbon layer with acetic acid. J. VÁČKA AND J. VAŠÁTKO. *Listy Cukrovar.* 49, 171-5 (1930).—Activated charcoal which has the slowest filtration time shows the greatest no. of small grains and possesses the greatest surface and internal condensation of water vapors. With 10 g. of charcoal in a 50-cc. filter, H_2O at 20°, $p_n = 7.2$, and a head of pressure of 208 g. per sq. cm., the increasing order of filtration is: Supra-norit 5x, Supra-norit 3x, Carborafin, Radit, Supra-carborafin, Polycarbon, standard norit. The rate of flow of H_2O through the charcoals varies exponentially with the pressure. Changes in temp. alter the viscosity of H_2O . At higher temps. (85°), a rapid flow occurs during the first stage of filtration—the process of wetting the surfaces and packing of layers proceeds slowly at higher temps. The rate of flow through charcoal decreases with increased viscosity of sugar solns. The degree of coloring had no effect upon the rate of flow. Adsorption of CH_3COOH occurs very rapidly during the first addn. of CH_3COOH . Satn. was attained when 1 l. CH_3COOH (1818 g. per 100 cc.) was used per g. of charcoal. FRANK MARESH

The results of sampling raw sugar. JIKI VONDRÁK. *Listy Cukrovar.* 49, 175-80 (1930).—A survey of the result of using standard methods of prep. representative samples during the seasons of 1927-30 is reported. FRANK MARESH

The treatment of the after-liquors. G. DIERCK. *Listy Cukrovar.* 49, 181-2 (1930).—The diln. water was computed on the basis of non-sugars present in raw sugar juices and added while the digested liquor was being released. During the outflow, an analysis was made, and the final diln. was made during the entrance into refrigerators. Samples of computations are given. FRANK MARESH

The course of the first saturation. III. Reproducibility of the saturation with lime and sugar solutions. J. DĚŘEK. *Listy Cukrovar.* 49, 183-90 (1930).—App. used in previous expts. were not capable of maintaining identical conditions throughout the course of the whole expt. The accuracy of the previous observations and conclusions is lowered. Further observation of the satn. process in plants showed that it is more complicated than lab. findings indicated. Using a highly refined app. with precise controls, D. ran over 100 expts. with pure sucrose and CaO from marble and observed total alk. alk. of the filtrate, absorption of CO_2 , and elec. cond. Detailed curves are thoroughly discussed. Although the curves show great similarity, D. was not able to attain an abs. reproduction under the refined method used, especially in elec. cond. Curves of changes of alk. at 20°, 50°, 70°, 90° do not resemble those of elec. cond. at the same temp., at higher temps. elec. cond. curves show 2 characteristic transformation maxima. In some instances the elec. cond. continued to diminish markedly even after abs. neutrality had been attained. This is due to a delayed hydration of CO_2 and changes accompanying the formation of sediment. The $CaCO_3$ comes out of soln. in a "firm form" more slowly than when it is formed by neutralizing CaO with CO_2 in satd. or supersatd. solns. IV. The union of lime in sediments. K. KOBLIHA AND

E. PALEČEK. *Ibid* 190-203.—The suspensions of CaCO_3 or CaC_2O_4 which form in a sucrose- CaO soln retain some of the free CaO . The concn of CaO in the sediment depends upon the alkyl of the filtrate, s , upon the actual concn of CaO in soln. This dependence has the form of an adsorption isotherm. The removal of CaO from the sediment is very rapid during the first 0.25-0.5 hr., the remainder is not released completely during the next 48 hrs. K. assumes that the CaO in suspensions is in 2 states: (1) CaO which is immediately sol., (2) CaO which becomes sol. after the first is dissolved. When oxalates were suspended in a CaO -sucrose soln, CaO was removed from soln.

F. M.

How much non-sugar from the water used for diffusion passes into the juice [in beet-sugar factories]? V. STANĚK AND J. VONDRAK. *Z. Zuckerind. czechoslovak Rep.* 54, 117-21 (1929).—The water supplies of some beet-sugar factories are liable, at times, to contain abnormal amts. of dissolved solids. To ascertain what proportion of the dissolved solids in the water passes into the juice the authors carried out comparative diffusion expts. with a lab. battery, using distilled water alone and distilled water contg. 0.1 and 0.5% of NaCl and urea. From the amts. of these substances found in the diffusion juices it is concluded that about $\frac{1}{3}$ of the dissolved matter in the water passes into the juice. If this is not removed by carbonatation, it must lower the purity of the final juice and increase the amt. of molasses obtained. The extra molasses due to 0.1% of mineral salts or 0.1% of org. non-sugar in the original water may amount to 0.25% of the wt. of beets in the former case or 0.12% in the latter.

B. C. A.

The influence of invert sugar upon the removal of amino acids from sugar solutions during saturation. VÁCLAV KOHN. *Listy Cukrovar.* 49, 194-5 (1930). The presence of a mixt. of glucose and fructose decreased the amino N removed with the sediment. Controls made without the invert sugar but with increasing CaO content also increased the amino N in the sediment. Glutamic and aspartic acids, asparagine and leucine were studied. The removal of the amino acids depends upon the presence of physically active CaO , s , lime which forms CaCO_3 during satn. and carries the amino acids into the sediment by adsorption.

FRANK MARESH

The balance of analysis of beet juices. FERDINAND KAYL. *Listy Cukrovar.* 49, 195-8 (1930).—Pressed juice from beet roots is not identical with the natural beet juice; the pressed juice contains natural beet juice secreted by the cells and admixts. of juices from cell walls and interstitial tissues.

FRANK MARESH

Scale from (sugar-house) evaporators. E. SAILLARD. *Circ. hebdom. comité central fabr. sucre* 42, 627 (1930); *Chimie & Industrie* 24, 1205-6 (1930).—Analyses were made of 93 samples of scale after a season's run. All contained CaC_2O_4 in variable amts., also insol. SiO_2 and sometimes Ca silicate , the latter being probably produced in the lime kiln because of excessive heating. There is generally 1-2% CaSO_4 originating from the limestone or from sulfitation. The amt. of CaCO_3 varies according to the efficiency of the boilers in decomp. and pptg. bicarbonates. The Fe_2O_3 and Al_2O_3 content decreases from the 1st effect (40%) to the last effect (about 1%), they originate chiefly from the limestone and coke. The points to be watched in order to reduce scaling are: the compn. of the limestone, the alkyl of the juices before boiling, the operation of the boilers and of the evaporators, the concn. of the wash liquors and the time of washing.

A. PAPIVEAU-COUTURE

The importance of a preliminary purification of the diffusion juice (beet). I. B. MINTZ AND B. E. KRASIL'SHIKOV. *Nauk. Zapiski Tsukrovoy Prom.* 9, 515-37 (1930).—Russian kieselguhr was compared with the best American and German grades and was found very much alike except for the wt. which is 3-4 times higher and the proportion of large particles is much greater. The filtration of diffusion juice through asbestos or fibrous material is advisable. The amt. of total N in carbonation juices, obtained from diffusion juices previously boiled or treated with 0.5-1.0% of kieselguhr, is 9-12% lower than that in ordinarily treated juices. The color of the juices at high alkyl is the same, at lower alkyl (0.02-0.01) the coloration of the preliminary treated juices is considerably lower than the color of juices treated in the usual way. Good clarification of the juice is obtained by preliminary sulfitation to a pH of 4.0 at a temp. of 20-40° followed by neutralization to alkyl of 0.1 with lime and increasing the temp. to 80°, then further filtration and carbonation to an alkyl of 0.01-0.02.

V. E. B.

Hyposulfites, redos, sulfoxylates and oxymethanesulfines. M. MESTRE. *Bull. assoc. chim. sucre dist.* 47, 388-93 (1930).—Based on his experience with hyposulfites, M. concludes that the Teatini process is sound as is indicated by the ultramicroscopic study of the colloids. The chemistry and application of sulfoxylates and hyposulfites are briefly reviewed. Oxymethanesulfine is entirely different from the others and is not used in the sugar industry.

F. CAMPS-CAMPINS

Return of greens and molasses into diffusion battery. I. B. MINTZ, G. S. BENIN, A. K. KARTASHOV AND B. E. KRAVITS'KHUKOV. *Nauk Zapiski Tsukrovos Prom* 9, 399-431(1930) — Expts on returning greens and molasses to diffusion juice continuously or intermittently proved unsuccessful. The total losses increased from 2.61 to 3.31%. The yield of final molasses increased and more fuel was used. The purity of the juices and sirups decreased. The expts were run in different beet sugar factories with practically the same results. V. E. BAIKOW

The investigation of the incrustation from the juice heater. I. O. TAEFIL'EV AND S. A. KOSTRO. *Nauk Zapiski Tsukrovos Prom* 9, 432-49(1930) V. E. B.

Investigation of the conditions for increasing the velocity of crystallization in "sugar loaves." M. I. NAKHMANDVICH AND I. F. ZELIKMAN. *Nauk Zapiski Tsukrovos Prom* 9, 450-92(1930) — A no. of expts in different refineries in Russia showed that velocity of crystn in sugar loaves greatly increases with a sharp decrease of the temp. The most convenient temp in the cooling rooms is $+5^{\circ}$ to $+10^{\circ}$. In summer the low temp can be maintained with the aid of ventilating fans. The increased rate of the cooling does not affect the quality of the refined sugar. The hardness and the color of the latter remain practically the same as during the cooling of the sugar loaves at the higher temps. The bleaching of the sugar is normal. A no. of tables and diagrams are included. V. E. BAIKOW

Boiling of a refinery massecuite with low-pressure steam. I. F. ZELIKMAN. *Nauk Zapiski Tsukrovos Prom* 9, 506-14(1930) — The excess of low pressure (1-2 atm) steam was used for boiling refinery massecuite. The increase in the time of boiling caused a slight increase in reducing sugars and coloring matters. These difficulties can be overcome by using a vacuum pan with a larger heating surface. V. E. BAIKOW

The bleaching effect of washing the massecuite of the second slap. V. V. CHESNICHENIK. *Zhur Sakharos Prom* 4, 303-8(1930) — The following expts were made to det the increase in yield and quality of white sugars resulting from the second slap. The massecuite was washed (A) with greens of 30° Bé, (B) with hot final molasses, and (C) with a measured quantity of molasses at $90-95^{\circ}$ (3.4-5.0% on the wt of massecuite) followed by warm water (1.8-2.0% on the wt. of massecuite). In (A) the yield of yellow sugar decreased 9.5% with purity increase of 0.9%, while color decreased 47.9% and the time of centrifuging decreased 19.2%. In (B) the yield of the yellow sugar decreased 1.5% with a purity increase of 0.9%, and the color decreased 23.1%. Purity of the molasses increased 0.5% and centrifugal time decreased 10%. In (C) the yield of the yellow sugar decreased 9.2%, while its purity increased 3.7%, color decreased 83.7% and the centrifugal time decreased 25%. Obviously method C is the most advantageous. V. F. BAIKOW

Cane sirup manufacture. J. O. CARRERO. *Porto Agr Expt Sta* 1929, Rept 13-4 (1930) — A light-colored sirup of full, agreeable flavor was manufd from sugar cane. The juice was readily clarified with H_2PO_4 and CaO but considerable sediment was formed during storage. The use of CaO always darkened the sirup. By restoring the cane juice to its original acidity after clarification, a better color was obtained. C. R. FELLERS

The beet crop in Czechoslovakia for 1930. JAROSLAV SOUCEK. *Listy Cukrotar* 49, 207-12(1930) — The sugar crop was above normal, the sugar content was normal, the total yield is estimated at 62.7×10^6 q of beets and 11.1×10^6 q of sugar. FRANK MARSH

A study of the biology of the beet-nematode, *Heterodera Schachtli* Schm. M. SINITSKII. *Nauk Zapiski Tsukrovos Prom* 9, 538-44(1930) V. E. B.

The physical and chemical properties of kaohang starch. I. The adsorption of iodine and chlorine by kaohang starch. SHOJI KIKI. *Abstract from Rept Central Lab S Manchuria Railway Co* 1929, 24-6 — The absorption of I and Cl is represented by the Freundlich equation $x = kx^a$, in which x is the quantity adsorbed and c is the concn of the Cl or I soln. A break in the curve of $\log c$ vs $\log x$ probably represents a change in surface tension due to swelling of the starch granule. V. F. HARRINGTON

Polarimetric determination of starch in potatoes and the relation between potato-dry substance and starch content. CARL VON SCHÖLKE. *Medd Kgl Lantbruksstyrelsen* 1930, No 283, 5-14 — Behrend, Maereker and Morgen (*Landw Vers Sta* 25, 1880) detd the relation between the sp gr of potatoes and the starch content, and prepd a table, which has not been used widely, and which contains errors. The relation between potato dry substance and starch content was found to be practically const at 5.75. The accuracy of the starch detn was $\pm 2\%$, attributed to air spaces. For starch detn 3 g of air-dried, finely ground potato heated with 50 cc of H_2O for

marinu Albatorokitra (Majunga). Tanning tests carried out on demineralized calfskin were quite satisfactory, showing that the bark approaches quebracho for tanning purposes, except that it gives somewhat too bright a reddish brown tinge to the leather. The bark would be suitable for the local production of dry exts., or for exportation for the production of liquid exts.

A PAPIEAU COUTURE
Tanning materials of New Zealand. P. WHITE. *Fall Imp Inst.* 25, 450-3 (1930).—*Tanekaha* (*Phoradendron trichomanoides*) bark was found to contain 17% tannins and 9.2% nontannins, as compared with 23-8% tannins reported by an earlier investigator probably because the later sample was from younger trees. *Kamahi* (*Hemmanthus racemosa*) bark from a 12 in diam tree, an 1½ in tree and a very old tree contained tannins 14.5, 22.2, 30.0 nontannins 5.9, 5.9, 6.4%, resp. *Tutu* (*Cordia ruscifolia*) has a very thin bark, which would not par for stripping, the combined wood and bark of a 1 year old shrub and of a 10-year old shrub had tannins 27, 31 nontannins 10.3, 8.3%. Samples of com black wattle (*Acacia mollissima*) bark grown in New Zealand contained 37-40% tannins and 9-11% nontannins; silver wattle (4 *dealbata*) bark contained 12.2-22.1% tannins and 9.5-12.1% nontannins, bark of a hybrid between black and silver wattle, which is immune to frosts and fungus disease contained 25.8% tannins and 11.3% nontannins. Bark from 7-, 16- and 26-year old trees of *A. decurrens* var *normalis* contained tannin 18.9, 28.4, 36.6, nontannins 13.1, 11.2, 10.8%, resp. Other species of acacia had the following tannin and nontannin contents: *A. melanoxylon* (Rotorua) 15.2, 9.9 *A. longistylis* 12.4, 8.0 *A. promensis* 23.4, 12.3, *A. Baileyana* 17.7, 10.1, *A. lunata* 22.4, 12.7, *A. decurrens* var *Leichardii* (New Plymouth) 33.3, 7.3, *A. pyramidalis* (New Plymouth) 28.2, 7.3%. Most of these samples were taken from comparatively old trees, viz., 20 yrs. Two samples of fungus galls (1 year old and much older) contained tannin 13.4, 28.3, nontannins 22.8, 9.2%, resp. The water ext. of the older galls was of a deep red color and had a pH of 5.9, on adjusting the latter to 4.8 the color was reduced nearly to that of a normal black wattle soln of similar strength. All results are given on a 10% water basis.

A PAPIEAU COUTURE
Skin glues and bone glues. H. MAIER BODE. *Kunststoffe u. Leim-Ind.* 26, 327-9, 330-41, 351-4 (1929).—It is confirmed that the viscosity of fresh glue liquor is smaller than that of a sol of equal concn. prepd. from the solid material, it is considered that, on drying, reaggregation of the gluten takes place. Skin glues and bone glues are conveniently tested by detg. the viscosity of a 17.75 or 20% soln. at 30° and 40°. A modification of Ostwald and Köhler's (C. A. 22, 2863) test differentiates skin glue from bone glue.

B C A

The condensation of catechol tannin (BERGMANN, POJARKIEFF) 10.

Bleaching leather, skins, furs, etc. ANDRÉ CORNILLON. *Fr* 692,981, July 6, 1929. Leather, skins, furs, etc., are bleached in a bath contg. peracetic acid prepd. *et al.*, by the action of H₂O₂ on Ag₂O. Salts such as phosphates, borates, alkali salts of org. acids or a base such as NH₃, are preferably present.

Greasing leather. HERMANN BOLLMANN and BRUNO REWALD. *Ger* 514,399, July 17, 1927. A mixt. of lecithin and fatty oil is used, *e.g.*, equal quantities of lecithin and train oil.

Leather substitutes. SOC. FLEURY ET LAVANANT. *Fr* 692,988, July 8, 1929. A leather substitute is made by treating unsized paper with strong H₂SO₄ charged with HCl applying the product to cloth and neutralizing, washing and drying.

Demineralization of bones. GÉLATINES HAASLT ET VILVOARE, SOC. ANON. Belg 371 121 July 31, 1930. The bones are treated with aq. H₃PO₄. The acid liquor is regenerated by adding it to a bath contg. a sufficient quantity of H₂SO₄ so that there will be no excess of liquor until the whole of the CaSO₄ has been pptd.

30—RUBBER AND ALLIED SUBSTANCES

C. C. DAVIS

The rubber industry in the United States. L. P. MAX. *Rev. gén. caoutchouc* 7, No 65, 3-6, No 66, 3-8 (1930).

C. C. DAVIS

A few facts on the study of rubber-bearing plants of the Russian Soviet Republic. SERGIUS IVANOV, T. S. CHELOVA, S. B. RESNIKOVA, E. I. TROFIMOVA, E. M. DUBNE AND V. I. SVESHNIKOVA. *Kaustschuk* 6, 237-9, 256-8 (1930).—The results of an in-

vestigation of over 150 plants are described, with analytical data tabulated in detail. The most promising raw material was the rubber of the *Chondrilla*, the "naphivi" of which with suitable improved methods yield considerable rubber C C DAVIS

The significance of the primary valence chain theory in explaining the structure of rubber. EUGENIO LINDMAYER *Kautschuk* 6, 249-52 (1930), cf C A 24, 750—A review and discussion C C DAVIS

The plasticizing of rubber. F JACONS. *Rev gén. caoutchouc* 7, No 67, 19-26 (1930), cf C A 25, 1117—Various wood resins, tars, pine tars and oils are described and discussed, with quant data (already published). C C DAVIS

An aging test with kerbosch rubber. R RIEM. *Arch Rubberculuur* 14, 411-20 (1930) (In briefer form in English 421-4)—Kerbosch rubber and latex sprayed rubber were kept at av temps of 27° and 18° (at places with different climates) to ascertain the influence of different climatic conditions on the aging of both types of whole latex rubber. The Kerbosch rubber deteriorated badly at sea level, but nowhere nearly so rapidly at 5000 ft elevation (cf de Vries, C A 18, 2265). The latex-sprayed rubber aged much better than the Kerbosch rubber, and furthermore the aging of the former was not so disproportionately bad at sea level. The results in general show the important influence of climate on the aging of rubber of the Kerbosch type C C D

Manufactured rubber, its keeping qualities and reconditioning. P BRÉRE *Arch méd pharm militaires* 91, 571 (1929), *J. pharm Alsace Lorraine* 57, 219-28 (1930), cf C A 24, 3397. S WALDBOTT

The Kelly abrasion machine. A. F HARDMAN, W. L MACKINNON AND S M JONES *Rubber Age* (N Y) 28, 463-5 (1931)—The new machine, which is described and illustrated, is designed to simulate the road wear of tires, with special attention to 3 factors: (1) increasing and diminishing pressure of contact of the rubber with the abrasive material, (2) application of only part of the surface to the abrasive material and (3) a side slip as the rubber leaves the abrasive surface. With this method, actual road wear can be foretold reliably C C DAVIS

Modern methods of cycle tire vulcanization. A. FRÖLICH *Gummi-Ztg* 45, 543 (1930), *India Rubber J.* 81, 154-7 (1931).—An illustrated description C C DAVIS

The influence of S dusting on rubber production (AMÉNT) 15. Rubberized cloth (Fr. pat. 693,890) 25. Rubberized conveyor belts (Brit pat 335,475) 1. SeS_2 [for vulcanization] (Fr. pat 692,758) 18.

VAN HARPEN, N H. The Electrometric Determination of the Hydrogen-Ion Concentration in the Latex of *Hevea Brasiliensis* and Its Applicability to Technical Problems. Medan (Java) Varekamp en Co 459 pp Reviewed in *Chimie & industrie* 25, 266 (1931).

Rubber articles from latex. SOC. ITALIANA PIRELLI. Brit. 335,621, June 27, 1929. In a process generally similar to that described in Brit 292,964 (C A 23, 1527), in which latex with a small proportion of certain added coagulants is subjected to local heating to produce coagulation, normal latex is heated to a moderate temp (suitably 40-60°) and cooled prior to the introduction of the coagulants. Various details of procedure are given.

Rubber. I. G. FARBENIND A-G (Curt Meisenburg, inventor) Ger 511,540, Nov. 13, 1925. Liquid polymerized substances are removed from rubber like condensation products (e. g. from butadiene) by treatment with powd absorbents such as active C, active SiO_2 , etc.

Rubber, etc., dispersions. THE DUNLOP RUBBER CO., LTD. and THE ANODE RUBBER CO. LTD. Fr 692,532, Mar 21, 1930. In the manuf of articles from org dispersions of rubber, gutta percha, balata or like vegetable resins, the surface of the molds or formers is covered with a film of dehydrating and coagulating liquid before immersion in the dispersions. Thus, the molds may be inserted in a concd soln of AcONH_4 , NaCl or NH_4Cl , to which an acid or a mixt. of acids may be added.

Rubber dispersions. NAUGATUCK CHEMICAL CO. Brit. 335,597, June 20, 1929. Rubber dispersions are prepd as water-in-oil emulsions by emulsifying water in oil with suitable emulsifying agents and adding rubber before, after or during the emulsification, with subsequent mixing to effect dispersion. The products can be used in substantially the same manner as rubber-solvent doughs are used in making spread or coated fabrics. Numerous details and examples of procedure are given. Cf C A 24, 751.

Preservation of rubber. WALDO L. SEMON (to The B. F. Goodrich Co.). Can

CHEMICAL ABSTRACTS

Vol. 25.

APRIL 20, 1931

No. 8

1—APPARATUS AND PLANT EQUIPMENT

W. L. BADGER

A device for determining work input to a laboratory ball mill. JOHN GROSS AND STUART R. ZIMMERLEY *Bur. Mines, Rept. of Investigations* 3056, 3 pp (1931).

ALDEN H. ENGRV

A satisfactory laboratory cement. V. T. JACKSON. *Chemist Analyst* 20, No. 2, 23 (1931) —Directions are given for making cements from asphalt, rosin, rubber and turpentine, such cements are used for making gas tight joints in apparatus. W. T. H.

The shape of clamps. WILHELM VOLKMAN. *Chem. Fabrik* 1931, 82-3, cf C. A. 23, 733 —The geometry of lab. clamps is discussed as a possible basis for standardization.

J. H. MOORE

Modification of the Seitz 20-ml. filter. N. W. LARKUM. *Am. J. Pub. Health* 21, 193-4 (1931), cf C. A. 23, 3601 —Wider and thicker flanges make possible the application of much more pressure, which is necessary to assure an adequate seal in order to prevent air leakage with its consequent contamination of the filtrate. J. A. KENNEDY

Filters for the reproduction of sunlight and daylight and the determination of color temperature. RAYMOND DAVIS AND K. S. GIBSON. *Bur. Standards, Miscellaneous Pub. No. 114*, 165 pp (1931).

E. H.

Mechanical stirrer. IRWIN STONE. *Chemist Analyst* 20, No. 2, 18-9 (1931) —Directions are given for making a stirrer; an old elec. motor or converted fan is used as power.

W. T. H.

A generally applicable shaking apparatus. P. DICKENS. *Chem. Fabrik* 1931, 61-2, cf C. A. 22, 3551, 23, 4905 —Two motor-driven shakers are described, one for use with ordinary glass app., the other for connected app. for carrying out reactions with air exclusion.

J. H. MOORE

A flask, divided in the middle, for evaporating and drying in a vacuum, boiling with reflux, and extracting. ARNOLD LUBINSKI. *Chem. Fabrik* 1931, 78.

J. H. MOORE

Improvised drying oven. GEO. W. CORRIE. *Chemist Analyst* 20, No. 2, 10 (1931) —For drying a no. of bulky objects at 50° it was found possible to make use of a small oven placed in the hood with a baffle plate in front of it. A cut of the arrangement is shown.

W. T. H.

Water baths. LESTER MEYER. *Chemist Analyst* 20, No. 2, 16 (1931) —An arrangement for keeping the level const. in the bath is shown.

W. T. H.

Hot water as a means of attaining higher temperatures. BATTIGE. *Apparatebau* 42, 217-19 (1930) —An air-heating system for use in drying rooms is described. Pipe bank "boiler" and heating coils.

M. C. ROGERS

Method for preventing an exhaust tube from shutting off on a water-cooled condenser. RUSSELL P. EASTON. *Chemist Analyst* 20, No. 2, 21 (1931) —By using a test-tube clamp on the upper outlet of a condenser mantle, collapse of the rubber tubing, which would impede the flow of the water, is prevented.

W. T. H.

A practical viscometer. H. R. HAMMOND. *Chemist Analyst* 20, No. 2, 21-2 (1931) —An improvised viscometer is shown which can be made at low cost.

W. T. H.

An adjustable picnometer float. FERDINAND AMBRUST. *Chemist Analyst* 20, No. 2, 22-3 (1931) —In detg. d by the picnometer method it is necessary to support the app. in a const. temp. bath so that it remains immersed to the proper depth. A floating support is shown which accomplishes this.

W. T. H.

Paraffining reagent bottles for alkaline solutions. WALTER B. DYER. *Chemist Analyst* 20, No. 2, 18 (1931) —Directions are given for lining a bottle with paraffin so that standard alkali soln. can be kept without deterioration as a result of action upon the glass.

W. T. H.

A modified Victor Meyer apparatus for the determination of molecular weights. EUGENE W. BLANK. *J. Chem. Education* 8, 546-7 (1931) —An easily constructed app.

is described which is accurate enough to give a mol wt of 76.5 for C_6H_6 when only about 0.02 g. is used
J. H. MOORE

Apparatus for melting-point determinations. M. SPETER. *Chem.-Ztg.* 54, 564 (1930) —The app. consists of a glass tube in the shape of a parallelogram with 2 sides vertical, contg H_2SO_4 acid or other suitable fluid. Near the upper end of the upper non vertical side is fused in a vertical tube of suitable size, into which the thermometer and m p tubes are placed so that their lower ends dip well into the circulating fluid, heat being applied at the diagonally opposite point. A sep tube of suitable length, closed at the upper end, serves as a cap to cover the thermometer, etc. B. C. A.

Apparatus for determining the fusion curves of low-boiling substances. SCHMOLKE. *Die Wärme* 54, 97-8 (1931) —App. designed by Simon, et al. (*C. A.* 23, 2640) for the detn. of the fusion curves of H, N, A and Ne is described. With the aid of this equipment and method, results were obtained with pressures in excess of 10,000 atm. Cf. *C. A.* 23, 5070 and also *C. A.* 24, 1559, 3411. ALLEN S. SMITH

Introduction of small, accurately weighed quantities of gas into evacuated receptacles. CARROLL W. GAFFIN. *Chemist Analyst* 20, No 2, 20-1 (1931) —The method and app. appear to be based upon that described by Pease (*C. A.* 17, 2220). W. T. H.

A broad source of monochromatic light. M. J. BURROUGHS AND V. F. HARRINGTON. *Am. Mineral* 15, 579-80 (1930), cf. *C. A.* 24, 1548 —A description is given of a remodeled Terrill burner equipped with a Pt gauze trough for holding the salt used to color the flame
A. M. BRANT

A new photoelectric photometer. B. LANGE. *Naturwissenschaften* 18, 917 (1930) —A priority claim (cf. Teichmann *C. A.* 23, 847) B. J. C. VAN DER LEEVEN

A recording dust-concentration meter and its application to the blast furnace. A. W. SIMON, L. C. KRON, C. H. WATSON AND H. RAYMOND. *Rev. Sci. Instruments* 2, 67-83 (1931) —The relation $\log(L_0/L) = k/m$, in which L_0 and L are, resp., the intensities of the incident and the emergent light, k is a const. characteristic of each dust, l is the distance traversed by the beam in the dust laden gas, and m is the mass of dust per unit vol., was theoretically deduced to connect the transmission of light and the concn. of a suspended material in a gas. To test the law, 2 sets of expts. were made. In one the concn. of suspended particles was kept const., while the distance through which the light beam passed was varied, in the other distance was kept const. and the concn. varied. While the derivation of the law was based on opaque material suspended in the gas where the grain size was large compared to the wave length of the radiation used, the tests showed that it also holds for white oil smoke. The app. consisted of an optical system, dust chamber, means for varying the length of beam and the concn., and a light-sensitive receiver. By fixing the length of path, the app. was adapted to installation in a blast furnace gas main. As a receiver, the thermopile was found to be superior to a photoelec. cell. The app. can be calibrated by comparison or by selecting an arbitrary unit of concn. Charts obtained under operating conditions show the use of the instrument adapted for recording instantaneous dust concns. A. S. S.

Fractionating head. H. MAINTZ. *Chem.-Ztg.* 54, 422 (1930) —The app. comprises a long wide tube with a side tube at the upper end for withdrawing the distillate, the whole being enclosed in a glass air jacket which serves to prevent temp. variations. The head is particularly useful for fractionating the carbolic oil fraction from tar distn.
B. C. A.

An inexpensive crystal or molasses separator. R. H. KING. *Ind. Eng. Chem.* 23, 300 (1931) —A filter made from pipe fittings for the sepn. of solids from viscous liquids is described.
L. W. T. CUMMINGS

Determination of the local strength of fabrics, felts, paper, rubber, films, wood, leather, metals, etc. I. FRIEDRICH SCHUREK. *Chem.-Ztg.* 55, 113-4 (1931) —The app. referred to in *C. A.* 23, 1187, 4103 has demonstrated its usefulness. It is described and its use and results are discussed.
J. H. MOORE

A design of experimental vacuum oven for temperatures less than 300°. G. FREDERICK SMITH AND V. R. HARDY. *J. Chem. Education* 8, 548-55 (1931) —A small tube oven, const. to $\approx 3"$ made of standard pipe and fittings, and heated by standard space heaters, is described in detail.
J. H. MOORE

The regulation of temperatures in industrial furnaces. V. PASCHKE. *Forschung Gebiete Ingenieurw.* Ausgabe A, 2, 29-40, 57-64 (1931) —The regulation of temp. difference depends upon the duration of the regulation period as well as upon the furnace and the regulator. The characteristics and sensitivity of the temp. regulation are discussed and exhibited in graphical form.
F. D. ROSSINI

The hydromechanical viewpoint in the construction of furnaces and boilers. F. MICHEL. *Feuerungstechnik* 19, 233-9 (1930) —Designers of boilers and furnaces have

not learned to apply modern discoveries in the field of fluid flow. Proper attention to these principles should make possible better utilization of the available combustion vol. reduction in friction losses and an improvement in heat transfer E. W. T.

Filter. DEUTSCHE LUFFTFILTER-BAUGESELLSCHAFT M. B. H. Ger 513,573, April 1, 1923 Details of a multistage gas, vapor and liquid filter are given.

Ultrafilter. STAATLICHE PORZELLAN-MANUFAKTUR. Ger. 513,861, June 12, 1923. The filter consists of balloon like cells of oval cross section, made of porous ceramic material. The cells are placed in a soln. and an elec. tension applied to the soln. via the filter cells.

Filter for liquids, with a number of filter elements supported in a rotatable frame. CHEM. FAB. GRÜNAU LANDSHOFF & MEYER A. G. (Walter Karchner, inventor). Ger. 516,459, Aug. 15, 1926.

Edge filters for liquids. SVENSKA AKKUMULATOR AKTIEBOLAGET JUNGNER. Brit. 336,122, Oct. 30, 1928. Structural features.

Apparatus for separating sediment, scum and gases from liquids. OSBERT DUNSCH. Ger. 516,439, June 30, 1928.

Air filter. I. L. BRAMWELL, C. W. H. HOLMES and BIRTLEY IRON CO., LTD. Brit. 335,922, March 26, 1929. Various features of a bag filter construction are described.

Filters for air or other gases. EDWARD L. JOSEPH. Fr. 694,024, April 16, 1930.

Centrifugal apparatus for separating suspended solids from gases. PNEUMATIC CONVEYANCE & EXTRACTION (1929), LTD., and W. A. SMITH. Brit. 336,620, June 17, 1929. Structural features.

Device for removing dust from air by centrifugal separation and straining. CLTDS C. MANCHESTER (to Tillotson Mfg. Co.). U. S. 1,791,732, Feb. 10. Structural features.

Device for removing dust from air in compressed-air lines. RICHARD M. FOX and HARVEY E. SMITH (to Acrometric Valve Corp.). U. S. 1,791,668, Feb. 10. Structural features.

Apparatus (with an oil- or water-moistened rotating disk) for removing dust from air. P. E. EGRNFELDT-OLSEN. Brit. 335,992, July 6, 1929. Structural features.

Filter presses. SOC. ANON. FOUR L'IND. CHIM. A. BALE. Fr. 694,717, April 29, 1930. A filter press is constructed with a heating element freely suspended inside the filtration chamber.

Fibers extracted from the coverings of coconuts by hot water for use as filters. LUDWIG THOMAS and ECON ELON. Ger. 513,765, Jan. 15, 1928.

Thermometer. Z. MATSUMOTO. Brit. 335,825, Nov. 29, 1929.

Thermometer of the variable electric-resistance and galvanometer type (suitable for use on internal-combustion engines). SAFE-T-STAT CO. Brit. 336,683, July 23, 1929. Structural and elec. features.

Temperature-indicating devices of fusible character for apparatus such as superheating or oil-cracking apparatus. CHARLES W. GORDON. (to Superheater Co.) U. S. 1,792,086, Feb. 10. Structural features.

Accelerator for temperature-control devices such as those of hot water-supply systems. JOHN M. LARSON (to National Regulator Co.). U. S. 1,792,230, Feb. 10. Structural features.

Radiation pyrometer. KAISER-WILHELM INSTITUT FÜR EISENFORSCHUNG. Brit. 336,296, Oct. 31, 1928. Structural and optical features.

Photometer for testing illumination of rooms, etc. HOLOPHANE, LTD., and H. S. ALLPRESS. Brit. 336,897, Jan. 20, 1930. Structural and optical features.

Viscometer. THE AUTOMATIC APPLIANCE CO. Fr. 694,376, April 24, 1930.

Calibrated viscometer. RICHARD VON DALLWITZ-WEGNER. Ger. 516,410, Mar. 17, 1926. Addn. to 515,259 (C. A. 25, 1415).

Röntgen-ray apparatus. SIEMENS-REISINGER-VEIHA GES. FÜR MEDIZINISCHE TECHNIK. Brit. 337,247, Jan. 4, 1929. Structural features.

Röntgen-ray apparatus. VICTOR X-RAY CORP. and BRITISH THOMSON-HOUSTON CO., LTD. Brit. 336,804, April 30, 1929. Structural features.

Röntgen-ray apparatus. VICTOR X-RAY CORP. and BRITISH THOMSON-HOUSTON CO., LTD. Brit. 336,953, July 18, 1929. Structural features of an app. adapted for stereoscopic work.

Röntgen-ray tube. C. H. F. MÜLLER A. G. Brit. 336,726, Aug. 27, 1929.

Vacuum tubes. E. Y. ROBINSON and ASSOCIATED ELECTRICAL INDUSTRIES, LTD. Brit. 336,710, Aug. 14, 1929. In applying insulating coatings to the heater wire of indirectly-heated thermionic cathodes by extrusion, as in the process of Brit. 320,866 (C. A. 24, 2336), the wire is first drawn through a plastic sludge of the insulating sub-

stance and a vehicle or agglutinant, and is then passed through a die before the coating is baked. App is described.

Photoelectric cell. S. RUNZ (to Aeternus Radio Tube Co.) Brit. 336,397, Jan. 11 1929. An electrode surface of etched Cu_2O is prep'd by heating a Cu electrode in an oxidizing atm at 1000° and immersing it in a soln of NH_4Cl or ZnCl_2 to produce a rough cryst surface. The electrode is embedded in beeswax and resin in a cell and exposed in an electrolyte such as NH_4Cl . Cf. C. A. 24, 2336.

Photoelectric cell. AUGUST KAROLUS and FRITZ SCHRÖTER (to Ges. für Drahtlose Telegraphie m. b. H.) U. S. 1,791,923, Feb. 10. Ne or He may be used as a filling.

Photoelectric cell. SOC. DES USINES CHIMIQUES RHÔNE-POULENC. Brit. 336,574, July 12, 1929. Various details are described of a gas filled cell which may have a cathode of a layer of K and an anode and auxiliary electrode in the form of straight parallel wires. The auxiliary electrode is near the anode and is maintained at a fixed potential much lower than that of the anode, so that it acts as a non photoelec. cathode when the potential applied to the anode nearly reaches the luminous discharge potential. Cf. C. A. 25, 238.

Electric discharge device. N.-V. PHILIPS' Gloeilampfabrieken. Brit. 336,563, July 8 1929. App for rectifying a c comprises a gas discharge tube in which are located, in sep. chambers, an incandescenl cathode (preferably an oxide cathode) and one or more anodes (such as graphite). Vaporized Hg may be condensed in the anode chamber. Various details of construction and operation are described. Cf. C. A. 25, 851.

Electric discharge tubes for emitting ultra-violet rays, etc. N.-V. PHILIPS' Gloeilampfabrieken. Brit. 336,222, July 5, 1929. For preventing impairment of the transparency of the tube walls to the rays generated, when the discharge is carried wholly or partly by a substance such as Mg which is likely to cause obscuration of the tube, the part of the tube wall intended for the passage of rays is coated with a substance (suitably KCl or KF in the case of Mg) which either prevents deposition or renders the deposit formed transparent to the rays desired. CaF_2 may be used as a coating for the interior of a Hg vapor lamp. K fluoroborate also is suitable for use with tubes contg. Mg. When the tube contains Na vapor, the wall may be coated with S or P. Various details are described.

Metallic-vapor luminous discharge tubes. PATENT-TREUHAND Ges. für ELEKTRISCHE GLÜHLAMPFEN (to General Elec. Co., Ltd.) Brit. 336,490, July 13, 1929. In a lamp with a filling comprising vapor of Na, K, Hg, Li, Ca or Zn, with a rare gas or a gas such as N to facilitate starting, condensation of vapor in the electrode chambers is prevented by encircling the part of the tube remote from the electrodes with laminated or ribbed metal rings to promote cooling. Oxide electrodes may be used, heated by a filament. Various details of construction are described.

Cathodes for electric discharge tubes. VEREINIGTE GLÜHLAMPEN- UND ELEKTRIZITÄTS- A. G. Austrian 120,121, June 15, 1930. Cathodes of elec. discharge tubes made wholly or partly of glass are oxidized *in situ* by electrolyzing the glass so as to liberate O within the tube. The glass is locally heated, and the low pressure gas filling of the tube is rendered conductive by maintaining a discharge in the tube. The cathode for the electrolysis is electrically connected outside the tube to the heated part of the glass, and the anode or cathode of the tube serves as the electrolytic anode. By interchanging the electrolytic anode and cathode, alkali or alk. earth metals may also be introduced into the tube. Cf. C. A. 25, 1125.

Catalytic hydrogenation apparatus. TECHNICAL RESEARCH WORKS, LTD., and E. J. LUSH. Brit. 336,569, July 11, 1929. In app having a Ni catalyst in the form of wire, turnings, or the like, packed in cages, the catalytic elements are constructed so that their thickness does not exceed 6 in., so that anodic oxidation may be thoroughly effected, an internal and an external cathode are used during the anodic oxidation process. Various details of construction and treatment are described.

Filler material for absorption towers. C. H. BUTCHER. Brit. 335,889, June 28, 1929. Hollow cylinders are used which after manuf. have had their sharp rough edges removed as by grinding, or tubes are used which are die pressed so that they are free from sharp edges.

"Molecular pump." E. R. GROTE. Brit. 336,001, July 9, 1929.

Mechanical-stage construction for microscopes. C. BECK and C. J. BECK. Brit. 336,064, Sept. 4, 1929.

Fine-focusing construction for microscopes. C. BECK and C. J. BECK. Brit. 336,063, Sept. 4, 1929.

Air-treatment, humidifying and dehumidifying apparatus. THOMAS CHESTER U. S. 1,791,751, Feb. 10 Structural features.

Apparatus (with a perforated inclined plate) for pneumatic separation of solids of different densities. R H KIRKUP Brit 336,055, Aug 29, 1929 Structural features

Continuous process for sterilizing milk or other liquids under pressure. N. J. NIELSEN Brit 337,027, June 27, 1929 Various details of app and procedure are described

Vacuum tank and heated tube system for condensing milk or other liquid products. N. J. NIELSEN Brit 335,904, June 5, 1929 Structural features and details of operation are described

Apparatus for pasteurizing milk or other liquids by the "Holding Process." RICHARD SELIGMAN (to Aluminium Plant & Vessel Co., Ltd) U. S. 1,792,328, Feb. 10 Structural features.

Apparatus for pasteurizing liquids in closed vessels. WILLY RÖCKH Ger 513,522, Nov 11, 1927. Details are given.

Apparatus for producing a foam-like mass from slag suitable for casting porous articles. NORSKE AKTIESELSKAB FOR ELEKTROKEMISK INDUSTRI Brit. 335,875, Feb 27, 1929 Structural features of an app with a horizontal or inclined container with stirrers, etc.

Apparatus and electrical system for magnetic tests of physical characteristics of iron or steel rods, etc. C KINSLEY Brit 336,987, June 19, 1929

Apparatus for continuous reactions between liquids under pressure. IMPERIAL CHEMICAL INDUSTRIES, LTD, and K H SANDRAS Brit 336,136, Nov 14, 1929 The app for carrying out continuous reactions between liquids under pressure described in Brit 329,200 (C. A. 24, 5188) is constructed of corrosion resisting material in the parts liable to be attacked by the reagents, e. g., in an app for the manuf of *p*-nitroaniline from *p*-chloronitrobenzene and aq. NH₃, the outer tube may be protected by a mantle of "stainless steel," and the central tube may be formed of the same material. Various details of construction and operation of the app are described.

Clarifying plant for liquids, particularly for coal and ore industries. PRÉPARATION INDUSTRIELLE DES COMBUSTIBLES. Ger 516,454, Feb 20, 1928.

Apparatus for determining specific gravity of solids or liquids. W. & T. AVERY, LTD, and W. A. BENTON Brit. 336,043, Aug 14, 1929 Numerous structural details are described

Sterilizing apparatus for liquids, with heated filter. CARPUL G M B II and OSWALD PONTICUS. Ger 516,007, Aug 19, 1927

Steam sterilizer. F & M. LAUTENSCHLAGER G M B II Ger. 516,008, May 23, 1928.

Mixing and emulsifying apparatus. VICTOR A COLLINS Ger 516,443, Oct. 3, 1929 See Brit. 323,534 (C. A. 24, 2920)

Apparatus for extracting gases from liquids such as oils. SOCIETÀ ITALIANA PIRELLI Fr 693,985, Apr 15, 1930

Apparatus for separating gaseous mixtures by liquefaction and rectification. SOCIÉTÉ L'AIR LIQUIDE (SOCIÉTÉ ANON POUR L'ÉTUDE ET L'EXPLOITATION DES PROCÉDÉS GEORGES CLAUDE) Fr 694,035, Apr 16, 1930

Apparatus for washing gases with atomized liquids. G. RAW. Brit. 336,221, July 4, 1929 Various details of construction and operation are described of apparatus suitable for removing dust from air or other gases

Air liquefaction system. M HAZARD-FLAMAND Brit 336,798, Oct. 18, 1928. Various details of app and procedure are described

Apparatus for storing liquefied gases. CHRISTIAN W. P. HEYLANDT. Fr 694,010, April 16, 1930.

Gas washers. PNEUMATIC CONVEYANCE & EXTRACTION (1929), LTD Ger. 513,527, Nov 15, 1928. Addn. to 508,571 (C. A. 25, 622) Further details of a partitioned air, etc., purifier are given.

Apparatus for drying gases with sulfuric acid. RENÉ HAYOT. Fr 694,038, April 16, 1930.

Wet and dry bulb apparatus for determining humidity of gases. CAMBRIDGE INSTRUMENT CO., LTD, J. L. ORCHARD and H T. GLOVER Brit 336,855, Dec 2, 1929 Various details are described of app employing elec resistance thermometers

Apparatus for effecting catalytic reactions between gases at raised temperature and pressure. MARIA CASALE-SACCHI Austrian 120,393, July 15, 1930.

Apparatus for ammonia synthesis and like catalytic gas reactions. HANS HARTER. Ger 516,249, Aug 11, 1928 See Brit 375,983 (C. A. 22, 2444).

Drying apparatus (spraying or atomizing type). PAUL KLAHR. Ger. 517,177, April 3, 1927

Drying apparatus with conveyors and several compartments separately controlled as to drying conditions. B. J. OWEN and R. O. DAVIES. Brit. 336,009, July 10, 1929 Various structural features and details of supply of heated drying gas, etc., are described

Apparatus for drying materials successively on superposed trays. L. N. LLOYD and PHILLIPS ENGINEERING Co., Ltd. Brit. 338,602, May 9, 1929 Structural features.

Tunnel drier and associated rotary air filter. TOMLINSONS (ROCHDALE), LTD., and J. N. TOMLINSON. Brit. 336,049, Aug. 22, 1929

Removing oil and fat from the hot chambers of drying apparatus. LOTHAR WOLF and KARL KUNZE. Ger. 514,074, Aug. 29, 1929 Addn. to 478,637 (C. A. 23, 4803).

Evaporating apparatus. Soc. DES CONDENSEURS DELAS. Ger. 516,365, May 22, 1927 Means is described for preventing the carrying away of froth by the vapors.

Apparatus for evaporating and cooling solutions. PAUL H. MCCLER. Ger. 513,789, Nov. 21, 1924.

Means for removing salts from evaporators operated at reduced pressure. PAUL H. MCCLER. Ger. 516,381, June 3, 1927

Apparatus for evaporating liquids or crystallizing salt solutions. E. M. S. INDUSTRIAL PROCESSES, LTD., RICHARD A. STOKES and EDWIN G. L. ROBERTS. Fr. 694,553, Apr. 25, 1930

Apparatus for developing crystals. IMPERIAL CHEMICAL INDUSTRIES, LTD. Fr. 694,014, Apr. 16, 1930 The vessel consists of an annular partition and means for circulating the liquor from the top downward in the central part and from the bottom upward on the sides.

Surface condenser suitable for use with steam. PAUL A. BANCEL (to Ingersoll-Rand Co.) U. S. 1,792,060, Feb. 10

Apparatus for cleaning steam. COCHRANE CORP. Ger. 516,158, Jan. 18, 1929

Apparatus for drying sludge for fertilizer or other materials by the action of heated circulating drying medium such as air or combustion products. JOHN B. BERRIGAN and JOHN J. BERRIGAN (to The Press and Drier Co.) U. S. 1,792,064, Feb. 10 Structural features.

Apparatus for eliminating carbon monoxide from motor-vehicle exhaust gases. HAROLD E. STORY. U. S. 1,791,912, Feb. 10 Various structural details are described of a device adapted for elec. heating of the gas in the presence of water vapor

Wet classifier for pulverulent materials. LEONARD ANDREWS. Ger. 516,146, July 30, 1927 See Brit. 280,121 (C. A. 22, 3005)

Heat-resisting cementation pot. FRIED KRUPP A.-G. Fr. 694,665, April 29, 1930.

Device for indicating the presence of coal gas, etc., in the air. RICHARD LEISER. Austrian 120,184, Feb. 15 1929

Horizontal chamber furnace. STETTINGE CHAMOTTE FABRIK A.-G. VORMALS DIDIER. Ger. 513,594, Mar. 1, 1928. Details of closure are given.

Door-lifting apparatus for chamber furnaces. EBERHARD GRASSHOFF. Ger. 513,571, April 1, 1927. Details are given.

Fusion furnaces. ITALSUC (SOCIETÀ ANONIMA). Fr. 694,186, April 18, 1930. Construction of burner for using powd. fuel is described.

Inclined grate furnace. L. and C. STEDNITZER. Ger. 513,476, May 13, 1928. Addn. to 511,566 (C. A. 25, 1126). Details of charging the forward part of the furnace.

Counterflow furnace suitable for heating articles in boxes. VEIT A. HAIN (to George J. Hagan Co.). U. S. 1,792,423, Feb. 10

Burner for liquid fuel. TOM E. A. NORDENSSON. Fr. 694,349, April 23, 1930.

Boiler plant fired with pulverulent fuel. INTERNATIONAL COMBUSTION ENGINEERING CORP. Ger. 516,449, Oct. 27, 1925

Pulverulent fuel burner. ROBERT S. WALKER. Ger. 516,257, Aug. 9, 1921.

Apparatus for distributing pulverulent fuel and air from a main to a number of branches. GEORGE R. ALDEN and CHARLES SKENTELBERT. Ger. 516,137, Oct. 7, 1928 Corresponds to Brit. 312,474.

Rotary-drum grate for finely granular fuels. WERNER SCHLEGEL. Ger. 516,483, Feb. 12, 1928.

Heat exchanger. SEMYON W. SYAKIN. Ger. 513,844, Mar. 25, 1928. A heat exchanger for pre-heating the fuel gas consists of a mass of tubes of elliptical cross section

Tubular heat-exchange apparatus. J. HOWDEN & Co., LTD., and J. H. HUMPHREY. Brit. 336,755, Sept. 21, 1929. Structural features of an app. with metal pieces having

- concave surfaces placed between rows of tubes to direct heating gases or the like in sinuous paths between the tubes.
- Tubular heat-exchange apparatus. C. A. HUNNAT. Brit. 337,236, Nov. 22, 1929
- Structural features
- Method of heating furnaces by the waste heat of spent gases. C. OTTO & Co G. M. & H. Ger 513,595, April 27, 1930
- Apparatus for removing ash without admitting air to the furnace. JULIUS JOSEF KELBER. Ger 513,792, July 17, 1928
- Gas burner. GASTON B. KILLAM U. S. 1,791,565, Feb. 10.
- Gas burner. CLARENCE H. MORROW (to Hotstream Heater Co.) U. S. 1,791,509, Feb. 10.
- Gas burner for heating boilers, furnaces, etc. J. E. WEYMAN Brit. 336,722, Aug. 23, 1929
- Safety valve for gas burners. CLARENCE H. MORROW (to Hotstream Heater Co.). U. S. 1,791,530, Feb. 10
- Structural details are described of a valve which cuts off gas supply to a main burner when a pilot light is extinguished
- Bunsen-burner attachment for producing radial horizontal flames. LOUIS TARTAZ. U. S. 1,791,913, Feb. 10
- Structural features
- Retort door for resisting pressure. NICHOLS PRODUCTS CORP. Brit. 335,946, July 3, 1929.
- Rotary kiln. HURON INDUSTRIES, INC., and B. E. GREEN Brit. 336,729, Aug. 23, 1929
- A sealing ring between a rotary kiln and its fixed ends is mounted so that its wt. tends to maintain a tight joint
- Tunnel kiln construction and heating system. V. LASTOVICKA. Brit. 336,171, Dec. 23, 1929
- Various structural details are described
- Electric water heater. L. P. THOMPSON. Brit. 336,011, July 11, 1929
- Structural features.
- Acetylene generator. AUTOGENWERK SIRIUS G. M. & H., THEODOR GEIS and AUGUST H. GIES Ger 513,612, May 16, 1929
- Details of app. for generating C_2H_2 from water and CaC_2 dust are given
- Acetylene generator and lamp. WM. P. HUNT and HESMANN C. STELLING (to Union Carbide Sales Co.). U. S. 1,791,595, Feb. 10.
- Thermostat. N.-V. NEDERLANDSCHE TECHNISCHE HANDEL MAATSCHAPPIJ "GIRO" Ger. 513,550, Sept. 22, 1929
- The device is applicable to hot water systems, the expansion or contraction of a rod opening or closing a rubber, etc., tube
- Thermostatic electric switch. R. MACLAREN Brit. 337,093 Aug. 10, 1929
- Thermostatic electric switch. FELTEN & GUTLEBAUM CARLSWERK A.-G. Brit. 337,162, May 6, 1929.
- Thermostatic device for electrically operated fire alarms. JAMES R. LONGWORTH and HERBERT HIRST (Longworth to Hirst) U. S. 1,791,598, Feb. 10
- Structural features
- Thermostatic and electrical control system suitable for piezoelectric crystal control systems. ALFRED CROSSLEY (to Federal Telegraph Co.). U. S. 1,791,804, Feb. 10
- Structural and elec. features
- Thermostatic cut-out device for electric circuits. ADONIRAM J. GREEN (to General Elec. Co.). U. S. 1,792,299, Feb. 10
- Structural features
- Thermostatic control for heating systems. JOHN M. LARSON (to National Regulator Co.). U. S. 1,792,101, Feb. 10
- Structural features of app. employing a plurality of thermostats.
- Thermostat construction suitable for control of fluid heating media. R. CARLSTEDT. Brit. 336,481, Nov. 25, 1929
- Thermostatically controlled reducing valve construction suitable for steam-heating systems. CLAYTON A. DUNHAM (to C. A. Dunham Co.). U. S. 1,792,213, Feb. 10.
- Thermostatically controlled combined hot-water supply and heating system. HOPE'S HEATING & LIGHTING, LTD., and J. H. HARPIN Brit. 336,125, Oct. 31, 1929.
- Structural and elec. control features.

2—GENERAL AND PHYSICAL CHEMISTRY

FREDERICK L. BROWNE

Life and work of the late Professor A. C. Baer. R. B. BECKER. *Proc. Assoc. Southern Agr. Workers, 31st Ann. Convention*, 158-62 (1930).—Baer's investigations on ice cream are discussed in particular.

K. D. JACOB

Two pioneers in agricultural chemistry of particular interest to the Association of Official Agricultural Chemistry. E. M. BAILEY *J Assoc Official Agr. Chem* 14, 18-28(1931) —An address on the careers of John Pitkin Norton and Samuel W. Johnson

A. PAPINEAU-COUTURE

Evgenii Vladislavovich Biron (E. W. Byron) II N. MENSCHUTKIN *J. Russ Phys. Chem Soc* 62, 1749-76(1930) —A biography with portrait

E. J. C.

Edward W. Morley, chemist, investigator, teacher CHARLES F. THIRING *Science* 73, 276-7(1931)

E. H.

Nikolai Aleksandrovich Shilov. S. VOZNESENSKIĖ *J Russ Phys.-Chem Soc* 62, 2102-12(1930) —A biography with portrait

E. J. C.

The activity of Professor F. Wald in analytical chemistry. O. QUADRAT *Collection Czechoslov Chem Comm* 3, 49-52(1931)

E. H.

F. Wald's theory of phases and of chemical stoichiometry. A. KKIĖ *Collection Czechoslov Chem Comm* 3, 9-31(1931)

E. H.

Bibliography of scientific communications [of Francis Wald] ANT. ŠIMEK *Collection Czechoslov Chem Comm* 3, 5-8(1931)

E. H.

The life of Professor Francis Wald. J. BABOVSKEĖ *Collection Czechoslov Chem Comm* 3, 3-4(1931). cf *C A* 25, 1128

E. H.

Harvey Washington Wiley (1844-1930) *J Assoc Official Agr Chem* 14, No 1, ii-xxi(Feb 15, 1931). The teacher. W. W. SKINNER iv-vi The chemist. C. A. BROWN vii-ix The leader. W. G. CAMPBELL ix-xi The pioneer. A. S. MITCHELL. xi-xiii The public servant. W. D. BICFLOW xiv-xv Cf *C A* 24, 5545

The organizer. II A. HUSTON xv-xvii The boss. MARY T. READ xviii-xix The man. T. D. LINTON xix-xxi

A. PAPINEAU-COUTURE

The relation of physics to chemistry. N. V. SIDGWICK *Science* 73, 209-76(1931)

E. H.

Rhodesian manganese mines of 7000 years ago. RAYMOND DENT *Can Mining J* 52, 180-1(1931) —The mines in the district east of Broken Hill, the primitive uses of minerals, the discovery of these mines and the uses of Mn in early civilizations are discussed

W. H. HOYNTON

The photoelectric measurement of luster. L. BLIV DPHLEDS *Dyer. Calico Printer* 64, 515-7, 583-5(1930). cf *C A* 24, 958

RUBY K. WÖRNER

Correlation of sieving analyses. II HAYWOOD *J Inst Fuel* 3, 428-32(1930) —The variable factors entering into a sieving analysis are: personal factor, nature of powder, time, weight of powder, variation in sieve aperture. These are discussed

D. A. REYNOLDS

Hydrographic sections and calculated currents in the Gulf of Alaska. GEORGE P. McEWEN, THOMAS G. THOMPSON AND RICHARD VAN CLEYVE *Rept of the Intern Fish Com. Commission No 4*, 5-36(1930) —[Printed by Wingley Printing Co., Vancouver, B. C., Can.] The water was classified into 3 arbitrary types: 1, coastal water with a chlorinity less than 18 per mille and a value σ_{θ} less than 25.50; intermediate water underlying the coastal waters and having a chlorinity between 18 per mille and 18.50 per mille and a value of σ_{θ} between 25.50 and 26.40; and ocean water with a chlorinity greater than 18.50 per mille and a value of σ_{θ} greater than 26.40. $\sigma_{\theta} = (d - 1) 10^3$, where the d is computed from the temp and chlorinity but is not corrected for the pressure p . In all 3 sections a cold surface layer lay over a deeper warmer stratum. The temp and chlorinity data were treated hydrodynamically according to Bjerknes' theory (Publ No 88 Carnegie Inst of Washington)

C. R. FILLERS

Mixture formulas K. K. JÄRVINEN *Acta Chem Fennica* 2, 80-6(1929) —Formulas are given for prep. certain solns and similar formulas in the handbooks are discussed. All formulas are derived from the equation (1) $k_1 p_1 + k_2 p_2 = k_3 p_3 = (k_1 + k_2) p_4$ where k means the wt in kg and p wts percent. In the following l = vol in l, v = vols percent or g in 100 cc or normality, s = sp gr, V = wt in vacuo and A = wt in air. By defining sp gr $s = k/l$ and substituting in formula

(1) (2) $l_1 s_1 p_1 + l_2 s_2 p_2 = l_3 s_3 p_3 = \frac{l_1 s_1 + l_2 s_2}{s_3} p_3$ Since $p = rd/s$ and $k = ls$ where d = density of the dissolved substance, one gets from (1) (3) $l_1 v_1 + l_2 v_2 = l_3 v_3 = \frac{l_1 s_1 + l_2 s_2}{s_3} v_3$

and for the contraction coeff K . (3a) $K = \frac{l_1}{l_1 + l_2} = \frac{l_1 s_1 + l_2 s_2}{(l_1 + l_2) s_3}$ If one puts

$p = rd/s$ in formula (1) (4) $\frac{k_1}{s_1} v_1 + \frac{k_2}{s_2} v_2 = \frac{k_3}{s_3} v_3 = \frac{(k_1 + k_2)}{s_3} v_3$ These formulas are

theoretically exact only for vacuum, metric system and $st/4V$ at a certain temp., but are practically exact also in other circumstances.

The atomic number of elements. A study of the specific gravity of elements in liquid form. L. H. BORGSTRÖM. *Finska Kemistsamfundets Meddelanden* 36, 70-81 (1927).

Was Mendeleev's system known prior to 1869? G. EISEN. *Chem. Weekblad* 27, 378-9 (1930)—A historical survey. The classification of the elements into 7 groups was proposed by Cooke in 1854.

Revision of the atomic weight of thallium. Analysis of thallous bromide. O. HÖNIGSCHMID AND H. STRAUFEL. *Z. anorg. allgem. Chem.* 194, 293-8 (1930)—The ratio $TlBr/Ag$ was detd. by gravimetric titration. The mean of 6 detns gave an at. wt. of 204.390 for Tl with an av. individual deviation of 0.008, confirming the previous result 204.39 derived from analysis of $TlCl$ (*C. A.* 18, 1593). The wt. of $AgBr$ formed by pptn. from $TlBr$ came out high, which was ascribed to occlusion, and the resulting detns. of the ratio $TlBr/AgBr$ were discarded.

An attempt to separate the isotopes of bromine. The atomic weight of bromine from the ratio of silver to silver bromide. P. HARTZKE AND H. STRIEBEL. *Z. anorg. allgem. Chem.* 194, 299-304 (1930)—The mass differences of isotopes are expected to cause slight differences in the sp. heats and consequently in the vapor pressures at low temps., on this basis a reversible sepn. is fundamentally possible. Fractional condensation of solid Br from the vapor in 3 stages (-18.5° , -25° , -36°) produced Br of at. wt. 79.916 \pm .001 in the most volatile fraction from an initial material of pure ordinary Br , at wt. 79.915 \pm .001, both detd. from the ratio $Ag/AgBr$. The results confirm the value 79.916 ($Ag = 107.84$) found by Hönigschmid and Zintl (*C. A.* 18, 771).

The interpretation of adsorption forces from atomic theory. F. LONDON AND M. POLANYI. *Naturwissenschaften* 18, 1099-1100 (1930)—Recently London (*C. A.* 24, 5597) advanced a theory of mol. cohesion on the basis of the action of virtual oscillators as known in the theory of light refraction. This same theory is applicable to adsorption. It is especially notable that this theory allows in first approximation simple superposition of the adsorption forces (cf. following abstract), this feature has not been included in other explanations of adsorption forces like the electrostatic theory or valence theory. The present theory is a good foundation for the potential theory of adsorption of Polanyi (*C. A.* 23, 5079). The adsorption potential of a mol. at a distance d from the surface is $\phi = -(N\pi\epsilon/6)(1/d^3)$, N being the no. of surface atoms per cm^2 , and ϵ a const. The const. can be calcd. from the dispersion curve of adsorbent and adsorbate, it consists mainly of the product $\alpha\alpha'$ of the 2 mol. polarizations. Elec. dipole moment and higher moments are generally only of secondary influence. The adsorption forces become of the same nature as the van der Waals cohesion forces, expressed in his const. α . Further conclusions are: Adsorption forces are largely independent of temp., if the mols. are not too large or the dipoles too strong. The forces are independent of other substances present on the adsorbent.

Properties and applications of molecular forces. F. LONDON. *Z. physik. Chem., Abt. B*, 11, 222-51 (1930)—The intermol. forces can be calcd. approx. from purely optical data (*C. A.* 24, 5597). In contradistinction to those observed electrostatically, forces so calcd. exhibit properties of a general cohesion, similar to the force of gravitation. The heat of sublimation, the heat of adsorption of adsorbed gases and the dissociation energy are calcd. from the same data. *Cf. C. A.* 23, 3399, 5094; 24, 2651. C. L. W.

Curie points. L. F. BATES. *Proc. Phys. Soc. London*, 43, Pt. I, 87-95 (1931)—A description of the magnetic behavior of a ferromagnetic substance may require 3 temps.: (1) the ferromagnetic crit. point, at which the rate of change of the square of the intrinsic magnetic moment per unit vol. with temp. is a max.; (2) the ferromagnetic Curie point, at which the intrinsic magnetization may be considered zero and (3) the paramagnetic Curie point, given by θ in the Curie-Weiss law $\chi = C/(T - \theta)$. The significance of the ferromagnetic and paramagnetic Curie points is discussed in the light of the results of a no. of investigators on Fe, Ni, Co, the ferrocobalts, magnetite and manganese arsenide. The paramagnetic behavior of these substances can be accounted for by a slight extension of the idea that ferromagnetism is due to a magnetic particle consisting of a group of assoc. atoms.

Magnetic susceptibility and chemical investigation. JAMES F. SPENCER. *J. Soc. Chem. Ind.* 50, 37-41T (1931).—Possible applications of measurements of magnetic susceptibility in various types of chem. research are outlined. Since magnetic susceptibility in the rare earths is additive, it can be used in fractionation as an index of the

degree of sepn. of the constituents. For this purpose it is often more convenient and more sensitive than the "mean equiv." In the study of alloys, variation of susceptibility with compn. often reveals intermetallic compds. This application is illustrated by examples from the work of a nn. of investigators on Pb-Sn, Bi-Sn, Pb-Ag, Pb-Au, and Al-Sn series of alloys. Evidence of O_2 mols. from susceptibility measurements is then discussed. Susceptibility detns. may be used to measure the velocities of chem. reactions and may also find application in studying assocn. and disocn. in binary mixts. of org. liquids. Calcns. based on susceptibility measurements prove that Willgerodt's paramagnetic org. compd. cannot have the formula usually attributed to it. Formation of the compd. BuO is highly improbable. W. W. STIFLER

Effect of internal stresses on the magnetic susceptibility of metals. KOTARO HONDA AND YOSOMATSU SHIMIZU. *Nature* 126, 900-1 (1930).—The adaptability of Cu to cold working was measured by Weiss's electromagnetic method. The decrease in d during cold working measures the internal stress. As the d decreased from 8.921 to 8.87, the susceptibility changed linearly from $\chi = -0.083 \times 10^{-6}$ to $\chi = +0.063 \times 10^{-6}$, that is, the Cu changed from diamagnetic to paramagnetic. On heating slowly to 500° the original diamagnetism was restored between 230° and 330° , proving that cold working set up internal stress. During cold working of Ag the d changed from 10.489 to 10.439 and the susceptibility from -0.200×10^{-6} to -0.140×10^{-6} . These results are in accord with H.'s theory of magnetism. W. W. STIFLER

Magnetic forces in a crystal of the type of rock salt. J. BOUMAN. *Arch. neerland. sci.* IIIA, 13, 1-28 (1930).—A highly math. discussion in which it is shown that certain crystals similar in type to rock salt may be constructed in which the atoms have definite magnetic moments. The assumption that the moments of the 2 atoms are equal is not necessary, although when this assumption is made the formulas are less complicated than those developed without it. W. W. STIFLER

Change of the dielectric constant of nitrobenzene with temperature. J. MAZUR. *Nature* 126, 993 (1930), cf. *C. A.* 25, 627.—The dielec. const. of carefully purified $PbNO_2$ was measured at a no. of temps. between $+30^\circ$ and -75° by a method based on the beat frequency between 2 high frequency circuits. The temps. were measured to 0.005° and the method was sufficiently sensitive to show a change in the 5th decimal place of the dielec. const. The dielec. const. increases steadily from 35.4 at 30° to 38.18 at 9.6° , then drops abruptly to 11.82 at 7.713° , after which it diminishes gradually to 9.707 at -75° . W. W. STIFLER

The thermal resistivity of solid dielectrics. ANON. *J. Inst. Elec. Eng. (London)* 68, 1313-55 (1930).—The study is divided into four parts: (1) a study of the thermal resistivity of solid dielectrics when tested in sheet form and of factors influencing the flow of heat through them; (2) a study of the thermal resistivity of solid dielectrics when applied to bars by the usual shop processes of wrapping; (3) a study of the thermal resistivity of sheet steel stampings transverse to the laminations and (4) a study of the thermal resistivity of air spaces. The appendix contains an estn. of the thermal conductivity of the insulation on a rectangular bar and a bibliography. The various forms of app. employed, as well as curves and results, are shown. W. H. BOYNTON

Specific heat of gases. H. LOURIE. *Chaleur et ind.* 11, 423-35 (1930), cf. *C. A.* 24, 5555.—A discussion of methods for detg. the length of sound waves in gases. The methods are classified as (1) dust methods, (2) methods of reinforcement of vibrations, (3) interference methods and (4) direct measurement. For the gases N_2 , O_2 , CO , air, CO_2 and H_2O at $0-2000^\circ$ the following are tabulated: (a) velocity of sound in an unlimited gaseous atm., (b) C_p at 1 kg., (c) C_v at 1 kg., (d) av. C_p at 1 kg., (e) av. C_v at 1 kg., (f) C_p at 1 cu m. (0/760), (g) C_v at 1 cu m. (0/760), (h) the resp. av. values. Values of numerous experimenters for O_2 , N_2 , H_2 , CO , CO_2 , air are also tabulated. Curves are given for the mol. sp. heats of these gases. 82 references. S. L. B. ETHERTON

Calculation of gas equilibria from spectroscopic data. L. J. M. BIRVOORT. *Chem. Weekblad* 28, 26-37 (1931).—A discussion of the calcn. of entropy consts. by statistical methods from spectral data. E. SCHOTTE

A measurement of molecular velocity and a test of the cosine law. P. CLAUSING. *Ann. Physik* [5], 7, 569-78 (1930).—Measurements of the time of passage t of A and Ne through a narrow cylindrical capillary at very low pressures, non-stationary flow being used, agreed within 1-3% with the formula $\bar{t} = L/\sqrt{4\pi u}$ (L = length, r = radius, u = $(8RT/\pi M)^{1/2}$ = mean mol. velocity). The derivation of this formula from kinetic theory assumes Maxwell's law for mol. velocities and the cosine law for the reflection (or reévapn. after temporary adsorption) of mols. from the walls. The accurate con-

firmation of the given formula for i excludes material deviations from the cosine law, e. g., an appreciable fraction of the molts specularly reflected. L. ONSAGER

Reflection and period of adherence of metal atoms on oil surfaces. G. VESZÍ *Z. physik. Chem., Abt. B*, 11, 211-21 (1930).—Vapors of Cd, Zn, Ti, Pb and Bi were directed through a narrow opening against a moving (15 m. per sec.) olive oil surface and reflected back a distance of 0.2 cm. onto a wall cooled by liquid air. The distance on both sides of the opening in which the thickness of the deposited metal was uniform was used as a measure of the distribution curve. This method gave definite proof that Zn, Cd and Bi adhere to the reflecting surface. The period of adherence was 10^{-4} to 10^{-6} sec. C. I. HOLST and CHIAUNG (C. A. 20, 1737) CURTIS L. WILSON

Vapor pressures of some hydrocarbons. ERNEST G. LINDER *J. Phys. Chem.* 35, 531-5 (1931).—The vapor pressures at temps. in the neighborhood of 0° are given for toluene, tetrahydronaphthalene, *m*-xylene, *o*-xylene, *p*-xylene, dipentene, ethylbenzene, mesitylene, octane, durene, tetradecane, decane, butylbenzene, *sec*-butylbenzene, *tert*-butylbenzene, *p*-diethylbenzene, *m*-diethylbenzene, propylbenzene, isopropylbenzene, octylene, diisobutylene, 2,2,4-trimethylpentane, decane, hexamethylthane, *p*-menthane, *p*-cymene, styrene, 1-methylcyclohexene, methylcyclohexene, limonene, pinene, decahydronaphthalene, methylnaphthalene, cyclohexene, diphenylmethane, acenaphthene and α -diphenylbenzene. MALCOLM DOLF

Heat transfer in liquid ammonia. III. M. HIRSCH *Z. ges. Kälte-Ind.* 38, 5-8 (1931) F. D. ROSSINI

Measurements with the aid of liquid helium. XI. The resistance of certain metals at low temperatures. W. MEISSNER and B. VOIGT *Ann. Physik* [5], 7, 701-97 (1930), cf. C. A. 25, 213.—The resistances of metals of the first 2 periodic groups and of Bi, Al, In and Tl were measured from 273°K. down to the temp. of liquid He. The characteristic temps. for these metals were called the Gruneisen formula being used. C. I. P. JEFFREYS

Determination of the coefficient of diffusion of metals in the solid state. SHINSUKE TANAKA and CHUJIRO MATANO *Proc. Phys. Math. Soc. Japan* [3], 12, 270-84 (1930); cf. C. A. 25, 1135.—The coeff. of diffusion was detd. at 500° for alternate layers of Au and Ag by measuring the change of elec. resistance with time. Fick's law was found valid. MALCOLM DOLF

The question of molecular or atomic solid solution of an intermetallic compound in a pure metal. G. WASSERMANN *Z. Metallkunde* 22, 158-60 (1930).—Soln. of Mg in Al increases a_1 (the side of the unit face-centered cube of Al) from 4.0104 A. U. to 4.0533 A. U. in an alloy contg. 3 at. % Mg. Soln. of Zn decreases a_1 from 4.0104 to 4.0330 A. U. in an alloy contg. 10 at. % Zn. When 1 at. % Mg and 2 at. % Zn (MgZn₂) are added to Al, a_1 increases to 4.0420 A. U., which is only slightly lower than the change in a_1 calculated additively from the effects of Mg and Zn taken separately. The vols. of 1 Mg atom plus 2 Zn atoms, and of 1 mol. of MgZn₂, are in the ratio of 1.08 to 0.98 on the basis of the vol. of 3 Al atoms (which they presumably replace in the solid soln.) = 1.00. Thus if MgZn₂ dissociates into atoms when dissolving in Al the Al lattice should expand, whereas if MgZn₂ dissolves as mols. the Al lattice should contract. The observed expansion of the Al lattice is taken as indicative of an at. solid soln. rather than a mol. ROBERT F. MEIN

The unipolarity of pressed lead sulfide. FR. TREY *Naturwissenschaften* 18, 1009 (1930).—PbS crystals with detector qualities become inactive after pulverization and pressing; pptd. PbS after being pressed to a pastille is also inactive. However, after heating to 150° the prepn. assume rectifier properties (C. A. 20, 1156). After repeated pressing these properties disappear again and reappear on heating. The resistance of the prepn. changes simultaneously with the rectification effect (Bridgman, C. A. 20, 1542). Both properties depend on the structural qualities of the crystals. B. J. C. VAN DER HOEVEN

Röntgenographic investigation of the system cadmium-magnesium. U. DENLINGER, *Z. anorg. allgem. Chem.* 194, 223-38 (1930).—In the system Cd-Mg there are 2 series of mixed crystals, α and β . Both are hexagonal. The α crystals exhibit the same axial ratio as pure Cd, namely $c/a = 1.89$, while β crystals, like pure Mg, exhibit the ratio $c/a = 1.62$. These ratios are independent of concn., and all lattices found belong to one of the types α and β . However, for crystals contg. 23-65 at. % of Mg, a transformation between the 2 may be effected by pressure (or mech. treatment), which favors β with approx. 2% smaller vol. Compds. with "hyperstructure" exist, the most important being MgCd₂ (α lattice) and CdMg₂ (β lattice), both were measured, the latter by the rotating-crystal method. A different type of hyperstructure was indicated by x-ray lines from alloys with about 50% Mg. L. ONSAGER

The ternary system potassium oxide, calcium oxide, silica. G. W. MOREY, F. C. KRACKER AND N. L. BOWEN. *J. Soc. Glass Tech.* 14, 142-87T (1939); cf. *C. A.* 24, 5443. —The phase relationships of these 3 rock forming ingredients are described in detail. The quenching method developed so successfully by the Geophysical Lab. was used, great care being observed to insure homogeneity of the melts. More compounds are formed in the $K_2O-CaO-SiO_2$ system than in the $Na_2O-CaO-SiO_2$ system, and there is a greater tendency for these compounds to decompose on melting, the viscosity of the melts at the liquidus is greater over a large proportion of the diagram, and the crystals and melting processes are more sluggish. The following new compounds were found and their equal fields and properties detailed:

	K ₂ O	Percent of CaO	SiO ₂	M. p. or Reaction	Crystal Habit
K ₂ O 2CaO 9SiO ₂	12.61	15.01	72.38	1070 = reaction	Prisms and plates
4K ₂ O CaO 10SiO ₂	26.46	5.43	58.12	946	Plates
2K ₂ O CaO 6SiO ₂	31.15	9.27	59.58	970	Plates
K ₂ O 2CaO 6SiO ₂	16.62	13.79	63.59	1115 reaction	Prisms
K ₂ O 3CaO 6SiO ₂	15.12	27.01	57.86	1020 reaction	Needles and plates
2K ₂ O CaO 7SiO ₂	41.36	13.20	42.43	1095 reaction	Octahedra
K ₂ O CaO SiO ₂	41.78	29.06	29.16	1020	Hexagonal bipyramids

	2 V	Optic Sign	Refractive Index	Crystal System
K ₂ O 2CaO 9SiO ₂	Large	Neg.	1.525 .. 1.515	Monoclinic or triclinic
4K ₂ O CaO 10SiO ₂	0	Neg.	1.551 .. 1.579	Hexagonal
2K ₂ O CaO 6SiO ₂	60 = 5.6-2	Neg.	1.543 1.541 1.525	Monoclinic or triclinic
K ₂ O 2CaO 6SiO ₂	Large	Pos.	1.50 ... 1.575	Orthorhombic
K ₂ O 3CaO 6SiO ₂		Neg.	1.57 1.56	Orthorhombic
2K ₂ O CaO 7SiO ₂			1.572 1.572	Pseudo-cubic
K ₂ O CaO SiO ₂	0	Pos.	1.605 ... 1.600	Hexagonal

H. F. L.

The crystal structure of cubic Carborundum. H. BRACKEN. *Z. Krist.* 75, 372-3(1939). —Last and rotation photographs on a small crystal of cubic SiC give results which agree with the powder photographs reported by Ott (*C. A.* 20, 3105).

L. S. RAMSDELL

The crystal structure of iron silicide, FeSi. V. WEVER AND H. MÜLLER. *Z. Krist.* 75, 3-2-5(1939). —A new determination of the structure of FeSi (cf. *C. A.* 13, 3153) gives $a = 4.47$ A. U. There are 4 Fe atoms in the unit cell, and the Fe and Si atoms have the positions $4f$, with $x_f = 0.134$ and $x_{Si} = 0.844$. The distance Fe—Fe is 2.74 A. U. and Si—Si is 2.76 A. U. Since each atom has a single atom of the opposite kind as its nearest neighbor, the structure can be considered as made up of FeSi pairs.

L. S. RAMSDELL

Crystal structure of the compound Fe₃B. GUNAR HÄGG. *Z. physik. Chem.*, Abt. B, 11, 152-62(1939). —"A new structure is established for Fe₃B which, for spatial reasons, is more probable than the structure assumed by Wever and Müller (*C. A.* 25, 39-1). It also offers a better explanation of the observed distribution of intensities" (on powder photographs). "The Fe atoms are located in the corners of tetrahedrons that form a highly uniform system. The B atoms are located in the largest intervals between these tetrahedrons." "Each B atom is surrounded, at equal distances of 2.176 A. U. by 8 Fe atoms. This structure indicates that the B atom has a radius of 0.97 A. U."

G. TÖRNQVIST

The cubic high-temperature structure of some perchlorates. H. BRACKEN AND L. HARANG. *Z. Krist.* 75, 323-43(1939). —The structure details by Herrmann and Dye (*C. A.* 25, 1429) for the alkali perchlorates are largely verified, although there are some minor discrepancies. Parameters and interatomic distances are given.

L. S. RAMSDELL

The crystal structure of calcium chromate. J. H. CLOVER. *Z. Krist.* 76, 285-6(1939)(in English). —X-ray data show definitely that CaCrO₄ is not orthorhombic and isomorphous with CaSO₄, but rather is tetragonal, and isomorphous with zircon, ZrSiO₄. The unit cell contains 4 molecules, $a = 7.10$ A. U., $c = 6.19$ A. U., the calculated d is 3.27; the space group is D_{2h}^{12} .

L. S. RAMSDELL

Large transparent cubic crystals of sodium chloride. TAKEMARO YAMAMOTO.

Bull. Inst Phys Chem Research (Tokyo) 10, 52-60(1931)(Abstracts 4-5 (in Esperanto) published with *Sci Papers Inst Phys Chem Research (Tokyo)* 15, Nos 283-5) —Various cations have a catalytic effect on the crystal by slow evapn of a satd NaCl soln. Mn^{++} gives large transparent crystals (the min concn is 0.01-0.001 g ion per l., by hydrolysis of $MnCl_2$ gives enough acidity but the optimum concn of H^+ is 0.001 N, the best temp is 40-60°). Pb^{++} acts similarly but the angles may be truncated (by 111). Sn^{++} and Cd^{++} give somewhat smaller transparent crystals. Nontransparent crystals are given by the following cations, in the order of size of crystals: $Ni^{++} > Cu^{++}, Fe^{++}, Cr^{+++}, Ti^+, Rb^+ > Fe^{+++}, Hg^{++} > Al^{+++}$. AUSTIN M PATTERSON

Habit variation in crystals of barium and lead nitrates. HAROLD E BUCKLEY. *Z Krist* 76, 147-68(1930)(in English) —The erratic choice of habit when $Ba(NO_3)_2$ crystallizes from soln at const temp is traced to the effect of humidity on the rate of evapn and to traces of $BaCO_3$. For $Ba(NO_3)_2$, the cube planes are increased (relatively to 111) by addn of carbonates of Ba, Na, Ca, Mg and Fe, by $FeCl_3$, methylene blue, Ba or Na manganate, Ba, Na or K permanganate, urea and Congo red. Cube planes are decreased by HCl , Ba, Na or K chlorate and perchlorate, safranin and quinine. The 210 planes are increased by $K_2Fe(CN)_6$ and $K_3Fe(CN)_6$, K and Ba chromate and $K_2Cr_2O_7$. For $Pb(NO_3)_2$, the cube is increased by Congo red, safranin, methylene blue, the cube is decreased by Na and K chlorates. Various theories concerning crystal habit are discussed. Observations on a large no of crystals lead to the conclusion that the true symmetry of these nitrates is tetartohedral rather than pyritohedral. L S R

The refractive indexes of mixed crystals of magnesium sulfate and zinc sulfate. MARY W PORTER. *Z Krist* 75, 289-300(1930)(in English) —In a series of mixed crystals of $MgSO_4 \cdot 7H_2O$ and $ZnSO_4 \cdot 7H_2O$ the variations of the principal ns were found to be directly proportional to the compn, as expressed in vol or mol percentage, or in percentage by weight. L S RAMSDELL

The dependence of crystal spacing on crystal size. J E LEVYARD-JONES. *Z Krist* 75, 215-6(1930)(in English) —From the theoretical standpoint, an actual increase in lattice dimensions is to be expected for min crystals of the non ionic type. For a crystal of 500 atoms, the increase should be 5% for one 5 layers thick, each way, 7%, for one 3 layers thick, 14%. This is because the total attractive field decreases relatively to the total repulsive field when particle size is decreased. For ionic crystals the reverse is the case. L S RAMSDELL

Absorption of synthetic spinels colored by chromium and manganese. K. SCHLOSSMACHER. *Z Krist* 75, 399-409(1930), cf *C A* 24, 4482 —Absorption measurements with a photoelec. cell were made in the range 400 to 600 mμ on synthetic spinels ($MgO \cdot Al_2O_3 = 1:1$) colored by both Cr and Mn. With relatively high Mn content the Mn absorption was superimposed on that of Cr. In the region 540 to 570 mμ there was absolute agreement with results obtained from a $KMnO_4$ soln, elsewhere there was only occasional agreement. L S RAMSDELL

A powder spectrometric study of urea. R. W. G. WATKOFF. *Z. Krist* 75, 529-37(1930) —Powder reflections combined with the best available estimates of at. F values give satisfactory results. It would appear that this method is useful for simple org crystals, instead of single-crystal spectrometry. For urea a unit cell was found with $a = 5.670$ Å. U, and $c = 4.726$ Å. U. At positions are $u_x = 0.14$, $v_x = 0.32$, $w_x = 0.57-0.58$, $t_x = 0.17$. At positions can be fixed with as great accuracy as that provided by Laue and spectral photographic comparisons involving more complicated planes. L S RAMSDELL

A determination of the molecular weight of methylbixin by means of x-rays. J HENGSTENBERG AND R. KUHN. *Z Krist* 76, 174-8(1930) —The unit cell has the dimensions $a = 10.50$ Å. U, $b = 17.59$ Å. U, and $c = 13.4$ Å. U (≈ 0.05 Å. U). With 4 mols in this unit, the calcd mol wt. indicates the formula $C_{41}H_{48}O_8$ rather than $C_{40}H_{48}O_8$, as proposed by Karrer (*C A* 23, 4480). L S RAMSDELL

The crystal structure of the diphenylpolyenes. J. HENGSTENBERG AND R. KUHN. *Z Krist* 75, 301-10(1930) —The unit cells of the diphenylpolyenes are (monoclinic series) diphenylbutadiene $a = 7.71$ Å. U, $b = 11.70$ Å. U, $c = 13.41$ Å. U, $Z = 4$, diphenylhexatriene 6.33 Å. U, 7.43 Å. U, 14.43 Å. U, 2, diphenyloctatetraene 6.25 Å. U, 7.44 Å. U, 16.03 Å. U, 2, (orthorhombic series) diphenyldecapentaene 10.25 Å. U, 7.66 Å. U, 21.2 Å. U, 4, diphenyldodecabeptaene 10.20 Å. U, 7.60 Å. U, 23.58 Å. U, 4, diphenyltetradecaheptaene 10.2 Å. U, 7.57 Å. U, 25.95 Å. U, 4. In every case the mol wt calcd from x ray data agrees with that from the structural formula. In the hexatriene and octatetraene there are definite *trans trans* bindings. The mols have a plane oblong shape. L S RAMSDELL

The structure of quinoid compounds and of a molecular compound of the quin-

hydrone type. EDUARD HERTEL AND GEORG H. RÖMER. *Z. physik. Chem., Abt. B*, 11, 90-6(1930) — Anthraquinone of the rhombic system contains 8 mols in an elementary structure of only 3.95 Å U height. The ring planes in this structure are parallel but on different levels. In 2,7-dinitroanthraquinone (I) of the quadratic system, however, an elementary structure of 3.4 Å U height contains 4 mols, the ring planes of which must be arranged perpendicular to each other. The mol compd 1-fluorene (I 1, trichine) possesses an elementary structure the dimensions of which preclude a lattice structure according to the tridimensional alternation principle."

G. TOENNIES

Determination of the melting point of platinum. G. RIBAUD AND P. MOHR. *Compt. rend.* 192, 37 8(1931) — The method used was the tube method already described for Pd (C A 24, 1006) the technic being improved by the use of a thermoelectric pile. By optical extrapolation, the m. p. of Au being used, the m. p. of Pt was calcd. as 1762°. The error in this value is probably not more than 2°.

LOUISE KELLEY

The extent of dissociation of salts in water. II. Uni-bivalent salts. EDWYN C. RICHIELLO AND CYCIL W. DAVIES. *Trans. Faraday Soc.* 26, 592-600(1930), cf. C A 21, 3793 — Dissocn. was calcd. from cond. data by using the Onsager theory and assuming an intermediate ion. The values of K for the ions used are: $\text{CdCl}^+ 0.0101$, $\text{PbCl}^+ 0.0301$, $\text{CaNO}_3^+ 0.521$, $\text{SrNO}_3^+ 0.150$, $\text{BaNO}_3^+ 0.121$, $\text{CdNO}_3^+ 0.391$, $\text{PbNO}_3^+ 0.0647$, $\text{LiSO}_4^- 0.229$, $\text{NaSO}_4^- 0.194$, $\text{KSO}_4^- 0.151$, $\text{AgSO}_4^- 0.05$ and $\text{TlSO}_4^- 0.0472$. The transport numbers calcd. from these consts. agree with the exptl.

A. F.

Notes on the specific gravities of liquid and solid sulfuric acid. D. MCINTOSH. *Proc. Trans. Nova Scotian Inst. Sci.* 17, 270-61(1930) — H_2SO_4 from 60° to 66° B° can be kept almost indefinitely in solid CO_2 without crystn. The sp. gr. was measured in calibrated bulbs with graduated stems of capillary tubing.

60° B°	Temp	Sp. gr.	66° B°	Temp	Sp. gr.
Liquid	+15°	1.635	Liquid	+20°	1.838
	-40°	1.756		-78.5°	1.951
	-78.5°	1.801			
Solid	-43°	1.797	Solid	-78.5°	2.040
	-60°	1.811			
	-78.5°	1.838			

RACHEL BROWN

The chemical constant of hydrogen vapor and the entropy of crystalline hydrogen. T. E. STEAN. *Proc. Roy. Soc. (London)* A130, 367-70(1931) — From the Einstein-Bose statistics the vapor pressure of cryst. H at low temp. is calcd. The chem. consts. of H vapor at low and ordinary temps., resp., are -1.11 and -3.70. The effects upon equl. of differences between the moments of inertia of gaseous and cryst. H are investigated. The entropies of cryst. para-H, ortho-H and ordinary H at zero abs. are, resp., 0, $R \log 9$ and $\frac{1}{2}R \log 3 + R \log 4$.

G. M. MURPHY

The thermal properties of ammoniates and analogous compounds and their use in absorption refrigerating machines. R. PLANK AND L. VAHL. *Forsch. Gebiete Ingenieurw.*, Ausgabe A, 2, 11-8(1931) — Heats of reaction, heats of vaporization, vapor pressures and refrigerating capacities are discussed for the systems: (1) $\text{CaCl}_2 \cdot 8\text{NH}_3 \rightleftharpoons \text{CaCl}_2 \cdot 2\text{NH}_3 + 6\text{NH}_3$, (2) $\text{CaCl}_2 \cdot 6\text{CH}_3\text{NH}_2 \rightleftharpoons \text{CaCl}_2 \cdot 2\text{CH}_3\text{NH}_2 + 4\text{CH}_3\text{NH}_2$, (3) $\text{LiCl} \cdot 3\text{CH}_3\text{NH}_2 \rightleftharpoons \text{LiCl} + 3\text{CH}_3\text{NH}_2$, (4) $\text{LiCl} \cdot 2\text{CH}_3\text{NH}_2 \rightleftharpoons \text{LiCl} + 2\text{CH}_3\text{NH}_2$. The following new data are given for CH_3NH_2 vapor pressure, -77° to 9°, $\log p_{\text{mm}} = -\frac{2068.6839}{T} - 8.66437 \log T + 0.004785 T$, C_p , liquid, at 20° = 0.77 g.-cal. per g.

F. D. ROSSINI

Problems of present-day colloid chemistry. VII. H. R. KRUYT. *Chem. Weekblad* 27, 176-9(1930) — The phenomenon of lyotropy is discussed in the light of the differences of size of similar ions due to varying degrees of hydration. The greater the hydration of the ions, the less water is available for forming the stabilizing hydration layer of the colloidal particles, the conditions obtaining thus resemble those in a soln. of an org. substance, the soly of which is depressed by the addn. of electrolytes. The manner in which the water mols. are oriented around the colloidal particles is also of importance. VIII. *Ibid.* 241-4 — A survey of certain outstanding problems, including adsorption, thixotropy, gelatinization and the relationship of emulsions to true colloidal solns.

B. C. A.

Heats of adsorption and their bearing on the problem of adsorption. II. R. KRUYT AND JOHANNA G. MODDERMAN. *Chem. Reviews* 7, 259-346(1930) — A thorough review of the exptl. data on heats of adsorption is given. An equation is deduced for calcn. of

the isothermal heats of adsorption. Calcd and observed results agree fairly well. Complete agreement is not to be expected. The state of the adsorbed phase is discussed critically. Each of the equations now in use in connection with adsorption has a domain of usefulness but none of them expresses a complete isotherm at lower temps. Eight factors that influence adsorption are listed. One equation cannot include all of them. Many heats of adsorption are tabulated. Complete bibliography. L. H. R.

Adsorption of chromate ions by colloidal aluminum hydroxide. BEN H. PETERSON AND KEITH H. STOKES. *J. Phys. Chem.* 35, 649-652 (1931), cf. *C. A.* 24, 1008.—The adsorption of K_2CrO_4 by colloidal $Al(OH)_3$ was detd. at concns. above and below the pptn. value. Three types of adsorption curves result, representing (1) neutralization of the pos. charge on the suspended particle, (2) surface adsorption not elec. in type and (3) a combination of 1 and 2. It is not always permissible to extrapolate data to zero concn. of the adsorbed substance. HARRY B. WEISER.

Adsorption of titanium hydroxide sol. L. S. BRATIA AND S. GHOSH. *J. Indian Chem. Soc.* 7, 687-697 (1930).— $Ti(OH)_3$ sol was prepd. by dropping $TiCl_3$ slowly into water at 18° . The sol cannot be completely freed from HCl by dialysis. A sol contg. 15.2 g. TiO_2 per l. was found on prolonged dialysis to become more and more viscous and as soon as the acid concn. fell to pH 4.1 the whole mass coagulated as a gelatinous ppt. The sol adsorbed appreciable amts. of anions from solns. of Na and K salts. The order of adsorption expressed as g.-mol. per g. of adsorbent was found to be $Fe(CN)_6^{4-} > NO_3^- > SO_4^{2-} > Cl^- > S_2O_8^{2-} > Cr_2O_7^{2-} > Br^- > I^- > Fe(CN)_6^{3-} > IO_3^- > CrO_4^{2-} > OH^-$. The coagulating power of the anions was found to be in the following decreasing order $Fe(CN)_6^{4-} > Fe(CN)_6^{3-} > OH^- > SO_4^{2-} > Cr_2O_7^{2-} > CrO_4^{2-} > IO_3^- > S_2O_8^{2-} > NO_3^- > Cl^- > Br^-$. Strong adsorption does not always indicate high coagulating power in this sol. The adsorption of OH^- ions by this Ti sol was small. The high coagulating power of OH^- ions was due to the removal of the stabilizing H^+ ions. The sol adsorbed more anions than cations with the liberation of OH^- ions and a corresponding increase of pH . $Ti(OH)_3$ is more basic than acidic. On aging, the sol prepd. at 18° became stable toward electrolytes and showed a decrease in sp. cond. and viscosity. A sol prepd. at 55° did not change in stability on aging and the sp. cond. and viscosity increased very slightly. Adsorption of Ti^{4+} by the sol prepd. at 18° accounts for its stability. L. H. REYERSON.

Studies on electrokinetic potentials. VII. The temperature coefficient of the ζ -potential. HENRY B. BULL AND ROSS AIKEN GORTNER. *J. Phys. Chem.* 35, 450-466 (1931), cf. *C. A.* 25, 1138.—The temp. coeffs. of the ζ -potential for the interfaces water-cellulose, $1 \times 10^{-4} N$ NaCl-cellulose, and EtOH-cellulose were measured between 20° and 51° . The latter interface potential has a pos. temp. coeff., while the other 2 show a slight max. in the neighborhood of 40° . MALCOLM DOLE.

Structure of thin films formed from solutions of crystallizing and non-crystallizing substances. G. TAMMANN AND H. FLSNER VON GABOW. *Z. anorg. allgem. Chem.* 194, 268-72 (1930).—The surface films formed by evapn. of solns. contg. soaps or similar substances were studied by interference colors. Perrin (*C. A.* 24, 3411) found that the local thickness of a Na oleate film varies by steps of 4.2μ . This observation was confirmed, and the same formation of flows with step structure was observed for films of K undecylate and K myristate. Among a considerable no. of substances which were investigated, the phenomenon occurred only with Na and K salts of higher fatty acids. It is suggested that besides an elongated form of the mols., a film capable of swelling is required. L. ONSAGER.

Experiments with precipitated and colloidal manganese dioxide. JNANENDRANATH MUKHERJEE, SATYAPRASAD ROY CHOUDHURY AND M. R. SUBBA RAO. *J. Indian Chem. Soc.* 7, 803-13 (1930).—The results of adsorption, antagonism and coagulation expts. with both pos. and neg. MnO_2 sols and of adsorption expts. with negatively pptd. MnO_2 conflict in many respects with the findings of Dhar and his co-workers (*C. A.* 21, 3005, 22, 1075). Further study is needed to reveal the true chem. nature of the particles in these sols. OSCAR T. QUIMBY.

A critical study of Kohlschütter's method of preparing silver hydrosol. M. RAMAN NAYAR AND P. S. MACMAHON. *J. Indian Chem. Soc.* 7, 699-707 (1930).—The presence of alkali ions is not essential in the formation of a Kohlschütter sol as stated by Pauli, *C. A.* 18, 2629. The walls of the exptl. vessels det. the nature of the sol. Silvered vessels produce no sol or only thin, pale, white, unstable sols. Silica or glass vessels invariably give rise to wine red, stable sols. The contradiction between Pauli's results and those of N. and M. is traced to the neg. catalytic effect of the Ag vessels used by Pauli. Stable Kohlschütter sol was prepd. at 85° . Reduction of Ag_2O soln. near the b. p. produces a coarse suspension. L. H. REYERSON.

Regarding the colloidal nature of cuprammonium solution. ALFRED J. STAMM.

J Phys Chem 35, 650-60 (1931) — It was demonstrated by the aid of the ultracentrifuge that cuprammonium solns. contain colloidal material, probably $\text{Cu}(\text{OH})_2$. The sol is polydisperse, contg. particles ranging in diam. from 6.5 μ to more than 20.0 μ for the concns. investigated. HARRY B. WEISER

Micelles and base exchange. MERLE RANDALL AND JESSIE Y. CANN. *Chem. Reviews* 7, 369-405 (1930) — Micelles are considered intermediate steps between ordinary ions and solid crystals. The soln. of micelles is taken to be a homogeneous phase and the thermodynamics of such a system is therefore that of a homogeneous phase rather than that of a heterogeneous system. The theory of the formation of micelles is extended to very large aggregates. The electrolytic nature of cryst. and near cryst. micelles is considered and the activity coeff. of such cryst. micellar masses is shown to be practically zero. All substances forming neg. micelles in soln. should possess differential base-exchange properties. Base exchange in zeolites, the compn. of zeolites and the methods of prepg. artificial zeolites are discussed. In zeolites the neg. ions of the aluminosilicate are considered to form large micelles. The rapid base or acid exchange properties of solid solns. and those of zeolites are explained by the electrolytic nature of the solids. Base exchange is considered in other systems as well as in zeolites. L. H. REYERSON

Mechanism of the mutual coagulation process. HARRY B. WEISER AND THOMAS S. CHAPMAN. *J Phys Chem* 35, 543-56 (1931) — The zone of complete mutual coagulation of 2 sols. of opposite sign may be very narrow or quite broad. When a given series of pos. sols., e. g., is arranged in order of the optimum concn. for mutual coagulation on mixing with neg. sols., the order of the pos. sols. may vary widely with different neg. sols. The reason for this behavior is that the mutual pptg. power is not detd. exclusively by the charge on the particles. Other factors are (1) mutual adsorption of colloidal particles that is independent of the charge, (2) the presence of pptg. ions as impurities and (3) interaction between stabilizing ions. The latter factor is not a general cause of mutual coagulation but it may be important in certain cases. The investigations were carried out with the following sols. (a) positively charged Fe_2O_3 , Cr_2O_3 , CeO_2 , BaSO_4 . Night blue, (b) negatively charged SnO_2 , $\text{Zn}_2\text{Fe}(\text{CN})_6$, $\text{Cu}_2\text{Fe}(\text{CN})_6$. Congo red. As_2S_3 . H. B. W.

A contribution to the theory of thixotropy. E. A. HAUSER. *J Rheol* 2, 5-9 (1931) — The particles of a sol. added to a bentonite dispersion do not come into close contact with the particles of the latter but rebound while at some distance away. The effect is due to the formation of solvated hulls around every particle. The view that each mech. disturbance disrupts certain solvated hulls, thereby lessening the viscosity, is widely held but it is untenable, because there is no change in vol. On the other hand, thixotropy is characteristic of anisotropic particles. In extreme cases of scale like particles. The actual setting of the gel is an instantaneous effect of orientation and not a continuous change from liquid to solid. EUGENE C. BINGHAM

Lyophilic colloids. I. Hydration. S. M. LIPATOV AND L. N. KOROBKOVA. *Z. anorg. allgem. Chem.* 194, 369-76 (1930) — The substance used in these expts. is "geranic acid", $\text{C}_{20}\text{H}_{31}\text{N}_7\text{O}_8\text{S}_2$. In dil. aq. solns. it is slightly ionized, as shown by a small elec. cond. The mol. wt. found from diffusion expts. is 506, the theoretical value is 497. The mol. wt. calcd. from cryoscopic measurements varies with the concn. and is lower than the theoretical value. The failure of geranic acid to dialyze (nature of membrane not given) indicates hydration. The ability of "true" solns. of geranic acid to gelatinize at low temp. is further evidence of its hydration. Osmotic pressure detns. on solns. contg. 2-25 g. per l. show that 1 g. geranic acid holds 31 g. H_2O . Viscosity detns. on solns. contg. 0.175 to 1.40% geranic acid indicate a similar high hydration. Both the viscosity and the osmotic pressure of geranic acid dispersions are greatly lowered by the addn. of small quantities (0.25-1.0%) of tannic acid. II. Theory of macro- and micro-syneresis. *Ibid.* 377-82 — Geranic acid gels show a much larger syneresis in 0.1 N NaCl than in H_2O alone. The time required for the complete syneresis decreases as the concn. of geranic acid increases. The velocity of syneresis is a function of the distance between particles and depends upon the intensity of the attractive forces, which are the same in all directions. The velocity of syneresis follows the law of the first-order reaction. The concn. of geranic acid in the liquid squeezed out is independent of the gel concn. When the vol. of the disperse phase is just equal to the total vol. of the system, no syneresis is observed, the addn. of more geranic acid breaks up the structure of such a gel because of partial dehydration. The elec. charge is the only true stability factor of lyophilic systems. Hydration merely delays the changes. Also in *J. Russ. Phys. Chem. Soc.* 62, 1895-1905 (1930). OSCAR T. QUIMBY

Hysteresis in sol-gel transformations. S. N. BANERJI AND S. GHOSH. *Z. anorg.*

glacial AcOH or of other proteins in anhyd HCOOH , are miscible in all proportions and without flocculation with solns of cellulose acetate (I) or nitrate in AcOH; these acid proteocellulose solns remain stable during evapn. The membranes formed may be regarded as true solid solns, the structure of which depends on the ratio protein cellulose. A 10% soln of gelatin in cold glacial AcOH is first prepd. According to the percentage of gelatin desired in the membrane (the percentage may vary from 5 to 75), a suitable quantity of this soln is mixed with a 10% soln of I in AcOH. The mixt. is shaken energetically, and the membrane prepd as usual. If the drying is complete, a resistant membrane is obtained, the permeability of which depends on the condition of evapn of the original soln. Membranes with 5% gelatin are transparent; those with 25% are opalescent. The preceding method is applicable to the prepn of membranes contg other proteins, except that anhyd HCOOH must be used to dissolve the protein. A perfectly dry membrane contg 20 parts casein to 80 parts dry I is clear and does not give the Tyndall effect. Using the method described, a series of membranes contg egg albumin, gliadin and the proteins of serum were prepd. LOUISE KELLEY

Solubility of thallous iodate in ethyl alcohol-water mixtures. VICTOR K. LA MIA AND IRENE R. H. GOLDMAN. *J. Am. Chem. Soc.* 53, 473-6 (1931).—The application of Boro's equation for the electrostatic work of transfer of charged spheres of av. radius r was tested with TlIO_3 alc-water mixts. The soly of TlIO_3 was measured in EtOH-water mixts contg 0-47% EtOH. From the data, the av. radii were calcd compared with the values of the parameter "a" computed from soly. data in dil. salt solns.

ALLEN S. SMITH

Liquid ammonia as a solvent and the ammonia system of compounds. VII. The nature of free radicals, their preparation and properties, as revealed by studies in liquid ammonia solutions. WARREN C. JOHNSON AND W. CONRAD FRANKLIN. *J. Chem. Education* 7, 200-8 (1930). cf *C. A.* 24, 3486.—Kraus classified free radicals on the basis of their affinity for the electron into strongly electropos., strongly electroneg. and amphoteric. The strongly electropos. radicals include the tetraalkylammoniums, which behave like typical alkali metals, i. e., they give blue solns. in liquid NH_3 and form amalgams. The RIIg radicals exhibit metallic properties. The strongly electroneg. radicals include $(\text{CN})_2$, $(\text{OCN})_2$, $(\text{SCN})_2$, and $(\text{SCSN})_2$. They do not exist in the monomol. state and they behave chemically like the halogens. Among the amphoteric radicals are Ph_3C , Me_3Sn and Me_2Sn . These substances combine with both the strongly electropos. and strongly electroneg. elements. Liquid NH_3 is a suitable solvent for the prepn and study of the amphoteric free radicals because many of the substances from which the radicals are prepd. are sol. in liquid NH_3 , and do not react with it. Typical reactions in liquid NH_3 are $\text{Me}_3\text{SnCl} + 2\text{Na} \rightarrow \text{Me}_3\text{SnNa} + \text{NaCl}$, $\text{Me}_2\text{SnNa} + \text{Me}_3\text{SnCl} \rightarrow 2\text{Me}_2\text{Sn} + \text{NaCl}$, $\text{Me}_2\text{SnBr}_2 + 2\text{Na} \rightarrow 2\text{NaBr} + \text{Me}_2\text{Sn}$.

W. C. FRANKLIN

The connection between ionic mobility and diffusion velocity in solid salts. CARL WAGNER. *Z. physik. Chem., Abt. B*, 11, 139-51 (1930).—For mixed crystals of solid salts of $\text{AgI} + \text{CuI}$ or $\text{PbCl}_2 + \text{PbBr}_2$, a formula is derived to calc. the diffusion coeffs. of the ions from conductivities and transference nos. or from the migration velocities of the ions. From the conductivities of pure solid salts with exclusive cationic or anionic conductivities the self diffusion const. of the more mobile ion can be derived. The formulas used by Hevesy (*C. A.* 15, 3243-1440, 22, 1711, 23, 3395, 24, 1573) and by Tubandt, Reinhold and Jost (*C. A.* 22, 342, 23, 3612) are confirmed. It is assumed that the single ions move practically independently of each other (no common migration of cation and anion, no exchange of places between adjoining cations or anions). Approx. exptl. verification of the derived formula confirms this assumption. G. T.

The temperatures of maximum density and the surface tensions of aqueous solutions in relation to concentration. G. TAMMANY AND A. ROHMANN. *Z. anorg. allgem. Chem.* 194, 273-7 (1930), cf *C. A.* 23, 19.—The changes of the temp. of max. d_m and of the surface tension σ effected by addn. of different substances to water are discussed from the viewpoint of internal pressure. Solutes of high internal pressure cause a marked decrease of t_m , directly proportional to concn. Solute of low internal pressure cause a smaller decrease of t_m , occasionally preceded by an increase at low concns., and the decrease is always smaller (particularly at low concns.) than the amt. calcd. by assuming the thermal expansion curves of water and solute additive. The difference is attributed to the internal soln. pressure's (K) being smaller than that of water, which affords a basis for the calcn. of K . It is pointed out that $\Delta\sigma/\Delta x$ (x = concn.) has the same sign as $\Delta K/\Delta x$. For substances of high K , $\Delta\sigma/\Delta K$ ranges between 13 and 72, while for substances of low K , $\Delta\sigma/\Delta K$ takes values in the range 5000-70,000. This contrast is connected with the well known fact that substances lowering σ , which conc.

in the surface layer, effect a much greater abs. change than those that tend to keep away from the surface, raising η L. ONVAGER

Viscosity of electrolytes. W. E. JOY AND J. H. WOLFENDEN. *Nature* 126, 994-5 (1930) —The viscosities of dil. aq. solns. of KCl were measured at 18° in both silica and glass viscometers. The results in glass indicate viscosities greater than that of pure water for concns. up to 0.025 *N*. When extrapolated to zero concn. the curve agrees with the predictions of the Dole-Falkenhagen equation. The results with the silica viscometers are less concordant but agree in general with those with glass.

W. W. STIFLER

The measurement of electrolytic resistance by the barretter method. J. SCHIELE AND M. WIEN. *Ann. Physik* [5], 7, 624-32 (1930) —An oscillating circuit of 10,000-100,000 hertz is connected to a barretter for measuring resistance of electrolytic solns. The barretter is a d. c. bridge that is shielded from the high frequency circuit with choke coils and condensers. The effects of increase in the applied potential, of strong and of loose coupling, and of thin and of thick platinizing of the electrodes are shown. The limit of precision with the Kohlrausch method, using a telephone, is 30 sq. cm. ohm., by the barretter method, 1 sq. cm. ohm.

G. M. MURPHY

The measurement of the conductance of electrolytes. III. The design of cells. GRINNELL JONES AND GILES M. BOLLINGER. *J. Am. Chem. Soc.* 53, 411-51 (1931), cf. *C. A.* 23, 4613 —Data are given that confirm the observation of Parker (*C. A.* 17, 2382, 3440) that the ratio of the resistance of 2 conductance cells when filled with a common soln. is not independent of the sp. conductance of the soln. used. In addition to the error due to polarization, there is a capacitive shunt between parts of the cell of opposite polarity which makes the measured resistance too low. New cells are designed which give results reliable within 0.01%.

JOHN R. HILL

Conductivity measurements and titrations. A contribution to their technique. FRIEDRICH L. HAHN. *Z. Elektrochem.* 36, 989-91 (1930) —A Cu-Cu₂O-Pb detector is used as a rectifier and a galvanometer as a null instrument in cond. measurements. The method is fast and accurate and is especially adaptable to titrations.

H. F. J.

Reduction potential of quadrivalent to trivalent iridium in hydrochloric acid solution. SHO-CHOW WOO. *J. Am. Chem. Soc.* 53, 469-72 (1931) —The reduction potential at 25° is 1.021 v. for the reaction $\text{IrCl}_4^{4-} = \text{IrCl}_3^{3-} + e^-$. The solns. were 0.1 *M* in HCl and the concns. of the 2 ions of the reaction between 0.002 and 0.008 *f*. The free energy decrease is -23,840 cal. and the temp. coeff. between 25° and 20° is +0.0098 v. per degree, which makes $\Delta H = -30,400$ cal.

JOHN R. HILL

The electrolytic reduction of acid solutions of vanadium. F. FÖRSTER AND P. BÖTTCHER. *Z. physik. Chem., Abt. A*, 151, 321-400 (1930) —The purpose of this investigation was to find out to what degree the reduction of V solns. is similar to solns. of Mo. E. m. f. measurements were made on equal mixts. of V solns. of different states of oxidation at various acid concns., with electrodes of Pt, Ir and Au. On platinized electrodes the first stage of reduction runs smoothly, but the second stage encounters considerable polarization so that H₂ may be evolved. On bright Pt the first stage of reduction forms a diaphragm on the cathode, thereby producing polarization. As was the case with Mo this diaphragm consists of difficultly sol. products of hydrolysis. Au cathodes act similarly to Pt ones.

MALCOLM DOLE

An effect of the breadth of junction on the electromotive force of a simple concentration cell. GEO. SCATCHARD AND T. F. BUEHNER. *J. Am. Chem. Soc.* 53, 574-8 (1931) —The e. m. f. of HCl concn. cells increased a few hundredths of a mv. when the liquid junction was made very thin. Flowing junction electrodes were used. This increase is greater than would be calculated from the Debye-Hückel theory and is probably caused by the change in temp. in the boundary layer resulting from the heat of mixing of the 2 solns.

JOHN R. HILL

The quinhydrone electrode. I. J. LIVINGSTON R. MORGAN, OLIVE M. LAMBERT AND MARGARET A. CAMPBELL. *J. Am. Chem. Soc.* 53, 454-69 (1931) —A comprehensive study of the quinhydrone electrode was made to det. its reproducibility in 0.1 *N* HCl under different conditions. Flaws in metal electrodes were the sources of the greatest error. A method of making, cleansing and testing these which minimizes breakage is described. The effect of the size, compn., form, condition and age of the metal electrodes, with and without the use of Hg as a connecting medium, is reported. The purity and amt. of quinhydrone do not appreciably affect the results if the concn. is not too low. Stirring the solns. is recommended when feasible. II. OLIVE M. LAMBERT, J. LIVINGSTON R. MORGAN AND MARGARET A. CAMPBELL. *Ibid.* 597-604 —The effect of *N* on the reproducibility of the quinhydrone electrode in 0.1 *N* HCl was detd. Better results

were obtained when N was passed through the soln. Equil. was attained more rapidly with electrodes prep'd. with N than with air. F. W. MARSH

Effect of various manganese compounds on the quinhydrone electrode. S. OGORI AND H. KASHIMURA. *J. Sci. Sol. and Meas.* 3, 4(1922), *Proc. Intern. Soc. Sol. Sci.* 5, 104(1922).—The presence of oxides of Mn (MnO_2 and Mn_2O_3) in small proportions (0.01 and 0.05%) caused inaccuracies in pH measurements with the quinhydrone electrode. In solns known to contain Mn, however, no significant differences in pH values as det'd. by H and quinhydrone electrodes were observed. B. C. A.

The potential of the cadmium electrode. FREDERICK H. GETMAN. *J. Phys. Chem.* 35, 5-6(1931).—Single crystals of Cd function as exact and reproducible electrodes, but give potential values that are uniformly greater than electrodes of the polycryst. form. A revision of the normal electrode potential of Cd gives the value, $E_0 = 0.235$ v. MALCOLM DREW

The potential of passive iron. A. TRAVERS AND J. ARTHUR. *Compt. rend.* 192, 161-3(1931).—The e. m. f. of Fe in Na_2SO_4 soln. of different pH values and in various oxidizing agents was det'd. The values vary considerably with the conditions and time, finally becoming constant. There are degrees of passivity and the oxide-film theory is not generally applicable. V. F. HAZENBROOK

The passivity of chromium. L. EASON MÖLLER AND O. EASON. *Z. Elektrochem.* 35, 9-7-7-(1930).—Electrolytic Cr is passive at 6° in HCl , HF , HBr , H_2SO_4 , $HClO_4$, H_3PO_4 , and HNO_3 . In the first 4 of these acids the metal becomes active when the acid is warmed, and each acid has a definite characteristic activation temp. that is proportional to the diam. of its anion. In $HClO_4$ and H_3PO_4 the metal is not active at 10°. Passive Cr is made more active by removal of adsorbed H, but when placed in the active acids the potential gradually rises with time until it is greater than the reversible H potential, without the loss of its passivity by the Cr. When Cr is cathodically polarized in inactive acids (except H_3PO_4) it suddenly becomes active at a definite potential and the potential increases immediately to that of the Cr Cr^{++} equil. The passivity of Cr is due to a layer of oxide, the pores of which are permeable to certain anions. H. F. JENKINS

Atomic reactions that are affected by inertia. H. v. HARTER AND M. POLANYI. *Z. physik. Chem.*, Abt. B, 11, 97 (1930), cf. C. A. 22, 910.—The reaction of Na vapor with alkyl halides is accompanied by, at most, only a trace of luminescence. To show the dimensions of the Na alkyl halide flame a "Na 1" resonance lamp was used. Na_2 or H_2 was used as the inert gas for the combustion. The diam. of the flame is inversely proportional to the square root of the pressure of alkyl halide and at controlled pressures serves to measure the speed of reaction. With MeI reaction is practically instantaneous. With MeF it is slow and could not be measured. Series were also obtained with Na in Br_2 , the speed being greater than with MeI. The same relative order is obtained with $PbCl_2$, $PbBr_2$ and PbI_2 . Substitution of other halogens on a C increases the speed of reaction. Substitution of other groups for H increases the rate of reaction. The heats of activation of MeBr (3 cal.), MeCl (5 cal.), MeF (>20 cal.), can be correlated with Lenz's theory, but that of $(CH_3)_2$ cannot. The reaction of Na vapor with CCl_4 tends to form NaCl instead of NaCN with increase of temp. The free CH_3 liberated in these reactions with alkyl halides reacts with H_2 at 300-500° to give CH_4 and H. FORTIS D. SWELL

The reaction mechanism of the formation of hydrogen molecules from atoms. H. SCHÖFFER AND O. RICHENBERG. *Physik. Z.* 31, 9-0-1(1930), cf. C. A. 24, 2022, 4980.—If $a = a_1 + a_2 + (a_1/a_2)(a_2 = \text{no. of atoms, } a_1, a_2 = \text{constants})$ then a must be a linear function of a_2 , if the surface reaction can be suppressed. This was found to be the case. The values of a_1 at 50, 30 and 15 mm. Hg were det'd. and the conclusion was reached that space collisions between 2 atoms and 1 mol. preponderate. The at. collision diam. was found to be 3.5 times as great for triat. collisions as for collisions between 2 atoms and 1 mol. In the latter case, the collision diam. is only slightly larger than the gas-kinetic one. FRANK URBAN

Kinetics of the polymerization of ethylene at pressures above one atmosphere. POSEY N. PRASE. *J. Am. Chem. Soc.* 53, 613-9(1931).—The rate of polymerization of C_2H_4 was measured at 2.5, 5 and 10 atm. from 250° to 550°. Fifty % of the C_2H_4 may be polymerized to gaseous and liquid monoolenes without formation of more than 2% of H and sat'd. hydrocarbons. Butylene is concluded to be the primary reaction product. The reaction is of the second order and has an abnormally low temp. coeff. The mechanism of the autocat. reaction is discussed. S. LEVINE

Polymerization and thermal decomposition of acetylene. P. SCHLÖTZER AND M. EATON. *Helv. Chim. Acta* 13, 1127-31(1930).—Polymerization of C_2H_2 begins at

300°, but does not become rapid until 450°. Uncatalyzed spontaneous decompn. occurs only above 500°. Polymerization was carried out in glass tubes and the reaction velocity at various temps was measured by the decrease in pressure that accompanies polymerization. At 420° the pressure decreased 110 mm in 0.5 hr and 500 mm in 5 hrs., when gas analysis showed that 71.5% of the $C_{11}H_8$ had disappeared. Polymerization proceeds in the homogeneous phase and the reaction probably is bimol. Increased glass surface (10-100 times) has little effect on the reaction velocity, but catalysts such as Cu accelerate, while Fe powder or C (from $C_{11}H_8$) retards decrease in pressure. Polymerization products could not be isolated. The temps and pressures at which explosive spontaneous decompn. of $C_{11}H_8$ takes place were detd. Below 530°, no deflagration occurs at pressures under 3 atm; at higher temps it takes place under much lower pressures (as 700° and 0.6 atm). C, Fe and Fe oxide speed the relatively slow non-explosive decompn. but retard the explosive decompn. especially at higher temps. Although deflagration occurs as low as 590° at approx. 1000 mm in the absence of C and Fe, in their presence it does not occur at 1960 and 2400 mm, resp. Impurities usually present in crude $C_{11}H_8$ have no appreciable effect on the spontaneous decompn., but the presence of small quantities of O causes deflagration at much lower temps. With increasing amts of O, explosive decompn. occurs less readily and spontaneous combustion (oxidation) more readily. A. WIEBER

The thermal decomposition of nitrogen pentoxide at low pressures. J. H. HODGES AND E. F. LAMBERT. *Proc. Natl. Acad. Sci.* 17, 23-32 (1931).—Previous detns. of this rate gave conflicting results. The rate was detd. at 35° and at pressures from 0.12 to 0.001 mm. The reaction is unimol. above 0.06 mm partial pressure of N_2O_5 . Below this the rate decreases with the pressure and below 0.004 mm the reaction appears to be bimol. V. F. HARRINGTON

Equilibrium diagrams of the aluminum-manganese, copper-manganese and iron-manganese systems. TORAJIRO ISHIWARI. *Science Repts. Tôhoku Imp. Univ.* 19, 499-519 (1930).—By thermal analysis, elec. resistance measurements, dilatometric and magnetic analyses, new and corrected equl. diagrams of the Al-Mn, Cu-Mn and Fe-Mn binary systems are set forth. Fifty-six photomicrographs are reproduced. CURTIS L. WILSON

The equilibrium iron-iron carbide-oxygen. P. PINGAULT. *Compt. rend.* 192, 43-7 (1931).—The Fe-C-O diagram is completed by detg. the curve representing the equl. between Fe and Fe_3C in mixts. of CO and CO_2 . In the app. used the gas was passed continuously through a closed circuit. The 50-cc. samples were analyzed chemically with an accuracy of more than 0.5%. The Fe was used in the form of fine wire which made it possible with a small wt. of material to fill the whole tube and obtain a rather large reaction surface. The expts. were carried out at atm. pressure. To insure having no false equl. due to passive resistance, Fe was carburized in a mixt. rich in CO and decarburized in a mixt. rich in CO_2 . Analysis showed that the 2 methods gave the same results at the same temp. When a mixt. of Fe and C was kept in a limited quantity of O the results varied with the temp. above 740° Fe_3C was formed and CO_2 in the gaseous phase varied from 0 to 23% between 740° and 700°, the proportion of CO_2 was 23-12%, and Fe_3C was oxidized to Fe and CO_2 . Below 700° the content of CO_2 became sufficient to oxidize Fe. The following calcs. were made between 700° and 1000°, $CO_2 + Fe_3C \rightleftharpoons 3Fe + 2CO - 14.5 \text{ cal.}$, $3Fe + C \rightleftharpoons Fe_3C - 27.5 \text{ cal.}$. L. K.

Notes on catalysis. A. BEYER. *Tida* 8, 1287-97 (1930).—A general review of catalytic reactions, enumerated in chronological order, and of the various theories which have been proposed to explain catalysis. A. PAPINEAU-COUTURE

The heat capacities at low temperatures of manganese sulfide, ferrous sulfide and calcium sulfide. C. TRAVIS ANDERSON. *J. Am. Chem. Soc.* 53, 476-83 (1931).—Heat capacities were detd. for MnS , FeS and CaS between 60° and 300°K. The values obtained for entropy are compared with related thermal data found in the literature. R. H. LAMBERT

The heats of combustion of methane and carbon monoxide. FREDERICK D. ROSSINI. *Bur. Standards J. Research* 6, 37-49 (1931).—With the app. that was used in detg. the heat of formation of water, the heats of combustion of CH_4 and CO were measured. For 25° and a const. pressure of 1 atm., the values per mole are, resp., 890, -160 = 300 and 282,890 = 120 international joules, 212,790 = 70 and 67,623 = 30 g.-cal. F. D. ROSSINI

The heat of mixing in molten metals. MASUO KAWAKAMI. *Science Repts. Tôhoku Imp. Univ.* 19, 521-49 (1930).—A continuation of previous expts. (cf. C. A. 22, 718). With similar or slightly modified app. 25 binary systems were investigated. The at. heats of mixing for 12 systems having intermetallic compds. existing up to the

Liquidus were without exception pos and reached a max. at the compn. of the compd. as given. Sn-Mg (at 800°), value of Q_{max} = +3420 cal for 63.3 at. % Mg. Bi-Mg (800°), Q_{max} = +3050 cal, 53.4 at. % Mg. Pb-Mg (800°), Q_{max} = +2410 cal, 61.6 at. % Mg. Zn-Mg (800°), Q_{max} = +1520 cal, 39.5 at. % Mg; Cd-Sb (800°), Q_{max} = +829 cal, 46.7 at. % Sb. Mg-Al (800°), Q_{max} = +1130 cal, 47.8 at. % Al. Al-Ag (1050°), Q_{max} = +1242 cal, 74.1 at. % Ag. Mg-Ag (1050°), Q_{max} = +3050 cal, 46.1 at. % Ag. Sb-Ag (1050°), Q_{max} = +1192 cal, 71.6 at. % Ag. Sb-Cu (1200°), Q_{max} = +903 cal, 57.4 at. % Cu. Al-Cu (1200°), Q_{max} = +4880 cal, 62.2 at. % Cu. Sn-Cu (1200°), Q_{max} = +1077 cal, 71.4 at. % Cu. Because the heat for Sn-Cu was pos and a max at the concn Cu₅Sn, it was thought that this compd should persist up to the liquidus. Eight systems having no intermetallic compds. up to the liquidus showed neg heat effects. Bi-Sb (800°), Q_{max} = -309 cal, 48.6 at. % Sb. Sn-Al (800°), Q_{max} = -1630 cal, 59.8 at. % Al. Zn-Al (800°), Q_{max} = -1215 cal, 48.6 at. % Al. Bi-Ag (1050°), Q_{max} = -1078 cal, 59.8 at. % Ag. Pb-Ag (1050°), Q_{max} = -1680 cal, 60.5 at. % Ag. Pb-Cu (1200°), Q_{max} = -2030 cal, 62.3 at. % Cu; Bi-Cu (1200°), Q_{max} = -1650 cal, 56.4 at. % Cu. Ag-Cu (1200°), Q_{max} = -1105 cal, 50.3 at. % Cu. Three exceptions were found, of which 2, Pb-Sb (800°), Q_{max} = +107 cal, 45.3 at. % Sb, and Sn-Sb (800°), Q_{max} = +1182 cal, 55.7 at. % Sb, were attributed to the presence of Sb, which, like Bi (see C. A. 22, 718), has abnormal phys properties. In the other exception, Sn-Ag (1050°), Q_{max} = +1260 cal, 67.9 at. % Ag, the reliability of the equil diagram is questioned. The systems Cu-Au and Ag-Au showed very small heat changes and are offered as examples of ideal solns.

CURTIS L. WILSON

The relation of the viscosity of cellulose ester solutions and temperature (REAL, et al.) 23. The motion of ions and colloid particles in an electric field (PATRICK) 3. Magnetic susceptibility of certain complex Mo compounds (RAY) 3.

Anales de química y farmacia (Chile) (*New journal*). Published bimonthly at Casilla 3670, Santiago, Chile. Vol. I, No. 1 bears the date Jan-Feb, 1931. Price 10 pesos.

ALEXIS, L. V. Fundamentals in Physics and in Chemistry. New Orleans: Post-Modern Scientific Thought. 336 pp. Reviewed in *Chimie & industrie* 25, 265 (1931).

BAILEY, C. H. AND BAUSOR, H. W. Chemistry for Matriculation. 3rd ed. London: University Tutorial Press. 448 pp. 7s. 6d. Reviewed in *Chem. News* 142, 142 (1931).

BARR, GUY. A Monograph of Viscometry. London: Humphrey Milford. 318 pp. 30s.

BRUGGEMAN, D. A. G. Elastizitätskonstanten von Kristallaggregaten. Groningen: J. B. Wolters' Uitgevers-Maatschappij. N. V.

DARROW, F. L. The Story of Chemistry. New York: The Bobbs-Merrill Co. 528 pp. Reviewed in *Chimie & industrie* 25, 264 (1931).

DUCLAUX, JACQUES. Viscosité et rigidité des liquides. Paris: Soc. d'Edit. scientifiques. F. 9.

FRANKENBURGER, W. AND DÖRR, F. Katalyse. Berlin: Urban & Schwarzenberg. M. 3 50.

Glasnik Khemuskog Drustva Kraljevine Jugoslavije (Bulletin de la société chimique du royaume de Yougoslavie) (*New journal*). Published by the Chemical Society of the Kingdom of Yugoslavia at Topchich Venatz 21 a, Beograd, Jugoslavia. Vol. I, No. 1, appeared in 1930.

HEIL, J. W. Leerboek der Natuurkunde (Vloeistoffen, Gassen, Warmte). 6th ed., revised. Rotterdam: Nygh & van Ditmar. 434 pp. Fl. 4.50, bound, Fl. 5.25.

LANDOLT, HANS, AND BÖRNSTEIN, RICHARD. Physikalisch-chemische Tabellen. 5th ed., revised by Walter A. Roth and Karl Scheel. Erg.-Bd. II, Tl. 1. Berlin: J. Springer. 500 pp. Bound, M. 75.

LENOBLE, EMILE. Chimie générale. Les synthèses totales en chimie organique. Paris: Gauthier-Villars et Cie.

Les institutions d'enseignement supérieur et de recherches en Belgique. Brussels: Fonds National de la Recherche Scientifique. 150 pp. Reviewed in *Chimie & industrie* 25, 263 (1931).

NAUMANN, EDUAR. Einführung in die Bodenkunde der Seen. Stuttgart: L. Schweizerbart'sche Verlagsbuchhandlung, G. m. b. H. 126 pp. M. 16. Reviewed in *Nature* 127, 195 (1931).

OSTWALD, W.: *Kleines Praktikum der Kolloidchemie*. 7th ed., revised by P. Wolski and A. Kuhn. Dresden: T. Steinkopff. 174 pp. M. 3.60. Reviewed in *Chemistry & Industry* 50, 115(1931)

PLANCK, M.: *Thermodynamik*. 9th ed. Berlin and Leipzig: W. de Gruyter, 287 pp. Bound, M. 11.50. Reviewed in *Chemie & Industrie* 24, 1029(1930)

REMSEY-SREUBERT: *Einleitung in das Studium der Chemie*. 7th ed. Dresden: T. Steinkopff. 347 pp. Bound, M. 7.50

ROST, ERNST: *Chemie für Alle. Eine volkstümliche Einführung in die Chemie und ihre Anwendungen in die Technik*. Stuttgart: Dieck. 442 pp. M. 12; linen, M. 18.

SMITH HABERS: *Praktische Übungen zur Einführung in die Chemie*. 4th ed., edited by V. Kohlschütter. Karlsruhe: G. Braun. M. 5, bound, M. 5.60

WATERMAN, H. I., and GOUDRIAAN, F.: *Handleiding bij de praktische Oefeningen in het scheikundig Laboratorium, tevens beknopt Leerboek der kwalitatieve Analyse van anorganische Stoffen*. 3rd ed., revised. Dordrecht. G. van Herwijnen. 95 pp. Fl. 2.90

WHITEHEAD, A. N.: *La science et le monde moderne*. Paris: Dunod. 272 pp. F. 25

ZENNECK, J.: *Aus Physik und Technik. Vorträge und Aufsätze*. Stuttgart: F. Enke. 189 pp. M. 14, linen, M. 15.50

Measuring pH values. SOLOMON F. ACRE: *Fr* 694,938, May 1, 1930. A stable buffer standard of H ion concn. is composed of an acid alkali metal salt of a dibasic org. acid the anion and cation of which have equal ionic mobility, e. g. acid Na phthalate mixed with very pure water. An alkali metal carbonate and an indicator, the latter being preferably regulated so as to be isohydric with the acid Na phthalate, are also present.

3—SUBATOMIC PHENOMENA AND RADIOCHEMISTRY

W. ALBERT NOYES, JR.

Radioactivity and the thermal history of the earth. ARTHUR HOLMES. *Naturwissenschaften* 19, 73-9(1931).—A review with numerous references on the cooling process of the earth. B. J. C. VAN DER HOEVEN

Interaction of free electrons and radiation according to Dirac's theory of electrons and quantum-electrodynamics. I. TAMM. *Z. Physik* 62, 545-68(1930).—Math. The Klein-Nishina formula is deduced. Transitions between neg. energy levels are essential to the scattering formula. A simple method for calcn. of the wave function of a free electron is given. B. C. A.

Proper energy of an electron on the quantum theory of scattering. I. WALLER. *Z. Physik* 62, 673-6(1930), cf. *C. A.* 25, 23.—Math. The difficulty of an infinite proper energy of a free electron in an electromagnetic field is discussed with the use of Dirac's relativity and non relativity dynamics. B. C. A.

Mechanics of photons. V. FOCK. *Compt. rend.* 190, 1399-1401(1930).—A math. development is given of the mechanics of a photon by treatment of Maxwell's equations in a manner corresponding with the wave motion of a photon (light quantum). An equation for a single photon is deduced and extended to a number of photons by application of Dirac's method of "posterior" quantification. The reasoning may be applied to the deduction of Einstein's law of energy fluctuation and to the laws of the interaction of matter and light. B. C. A.

Oscillatory theory of quantum phenomena. L. DÉCOMBE. *Compt. rend.* 190, 1385-7(1930).—The author's theory (*C. A.* 24, 1575), when applied to band spectra, indicates that the Deslandres spectral term (*C. A.* 24, 3950) is attributable to ionic oscillations about the center of gravity of the mol. With the assumption of a const. H for the proton, analogous to the electronic quantum const. (h), it is shown that there is resonance between the 2 rotatory frequencies of the spinning proton and electron. B. C. A.

Application of group theory to the quantum dynamics of monatomic systems. CARL LICKART. *Rev. Modern Physics* 2, 305-80(1930).—The math. principles of finite group theory are discussed with special reference to quantum dynamics. The general theory includes discussion of invariant manifolds, Schur's lemmas, perturbation theory, multiplication of groups and orthogonality relations. The applications are concerned with complex at. spectra and include the method of sums, selection principles, alterna-

tion rule of multiplets, the permutation group and the Pauli principle. The influence of external fields is considered for the Zeeman and Stark effects. Mol spectra valence and electron conduction are not discussed. G M MURPHY

Magnetic susceptibility of certain complex molybdenum compounds. PRAYADA RANJAN RAY. *J Indian Chem Soc* 7, 741-3 (1930).—The mass susceptibilities of 5 complex coordination compds of quinquivalent Mo were measured at 29°. The results ranged from $\chi_m = 2.475 \times 10^{-4}$ for $(C_6H_5N)_3MoOCl_3$ to $\chi_m = 1.140 \times 10^{-4}$ for $(C_6H_5N)_3MoOBr_3$. The corresponding Weiss magneton nos are 8.073 and 7.257, and in every case the nearest Bohr magneton no is one. These results do not agree with Rose's rule for the magneton no and a table of possible electron arrangements to account for the anomaly is given. It is also suggested that the magnetic behavior of Mo may be masked by that of the org radical. W W. STURLEY

The motion of ions and colloid particles in an electric field. H H. PAINE. *South African J Sci* 27, 11-22 (1930).—An address, in which the analogy between ions and colloid particles as concerns their elec structure, is discussed from the standpoint of the Debye-Huckel theory. C E. P. JEFFREYS

The presence of hydrogen ions among the positive thermions as usually obtained. III. Catalytic hydrogenation. OTTO SCHUMER. *Z physik Chem, Abt A*, 152, 293-83 (1931).—Theoretically it appears impossible to form H^+ by thermal emission. S finds expl proof of this and so all theories of hydrogenation as an ionic reaction in the gas phase are invalid. This does not preclude the possibility of such reactions in the metal phase. The app is described and data are given. V P HARRINGTON

Measurements on the Langmuir dark space. A GENTHEASCHULZE. *Z Physik* 62, 819-23 (1930).—The Langmuir dark space which surrounds an electrode maintained at a neg potential, with respect to a region in which a discharge is maintained, is due to the relative inefficiency of pos ions in producing new ions on collision and is accompanied by a pos space charge. The dimensions of this dark space can be calcd, provided certain conditions are fulfilled, by the simple form of the space charge equation. Expts have been carried out in H, He, Ne and Ar in which the above mentioned conditions are approx fulfilled. A fairly satisfactory agreement exists between the observed and calcd dimensions of the dark space. B C. A.

Wave length of x-rays. T H LUBY and R BINGHAM. *Nature* 126, 915-6 (1930).—By means of a plane glass grating of 10 800 lines to the inch, a vacuum spectrograph and x ray tube with graphite target in the same vacuum, it was found that the carbon K line could be photographed from the 18th neg to the 17th pos order. With very small dispersion the carbon K line was compared with Cu L_{α} and L_{β} in their 2nd and 4th orders. The wave lengths 41.7 and 44.8 Å being obtained for the C line relative to 13.32 for the Cu L_{α} line. Al $K_{\alpha 1}$ = 8.315 Å U relative to Cu $K_{\alpha 1}$ = 1.5392 Å U. C E. P. JEFFREYS

Electron emission by collision of positive ions at low gas pressures. A GENTHEASCHULZE. *Z Physik* 62, 601-6 (1930).—The efficiency of pos ions for the liberation of electrons has been examd at a Cu surface acting as a subsidiary electrode adjacent to a low voltage arc in Ar, Ne and He, the pressure ranges being 0.0017-0.0036, 0.0120-0.0156 and 0.019-0.025 mm, resp. The efficiency is of the same order in all these gases and increases from about 10 to 30% as the potential of the subsidiary electrode is increased sevenfold. B C. A.

The emission of positive electricity from palladium. P. GILLES. *Acta Univ Latvianis, Mat un Dabaz Zinatnu Fakult*, Scripta I, Fasc. I, Burtinis, (in English 51-66) (in Lettish 67-8) (1931).—G investigates the pos (thermionic) emission of Pd in order to det. the decay of emission with time and the effect of occluded gases and dirty surfaces. The app used consisted of a loop of pure Pd wire surrounded by a cylinder of Pt foil, electrically connected to an electrometer. The 2 electrodes were enclosed in a pyrex bulb attached to a vacuum pump and an auxiliary spark gap for testing the reduced pressure. The Pd electrode was removable for cleaning. The values detd. for the decay of emission with time were not found to agree with the equation $i = i_0 e^{-t/\tau}$, but they more nearly agreed with the expression $i = K$. Attempts to revive the Pd by admission of dry air or H_2 for several hrs were unsuccessful, but contact with H_2SO_4 , HCl, H_2O , LiOH, Li_2O , grease and dust particles gave various increases in the emission. Immersion of the Pd wire in toluene or xylene caused no increase in emission. The theory is advanced that the pos emission of metals is due to some "active matter" which is widely distributed and which can be transferred from one body to another. The ions of the "active matter" are thought to be bound in the metal in such a way that a definite energy is required for their liberation. The process of emission is unknown but is maintained by itself and decays with time. Temp is only a promoter of it. H W. L.

Oxygen films on tungsten. I. A study of stability by means of electron emission in presence of cesium vapor. I. LANGMUIR AND D. S. VILLARS *J. Am. Chem. Soc.* 53, 486-87 (1931), cf. *C. A.* 19, 931 —W filaments in the presence of Cs vapor emit more electrons than the pure W filament. Up to 640°K at which emission is many billion-fold that of pure W the Richardson equation, $i = AT^2 e^{-b/T}$, is valid in the presence of Cs vapor. If O is admitted to activate the filament by forming a W-O-Cs surface the emission is greatly increased and the Richardson equation is valid to much higher temps. A theoretical discussion of results and a description of exptl. methods are given with suggestions for their application to the study of the loss of O from films on W and its detection in a gas. A. LEON TAYLOR

γ-Rays in the power field. ROBERT F. MEHL AND CHARLES S. BARRETT *Power* 73, 110-2 (1931) —The penetrative limit of γ rays through steel is about twice that of x rays. The expense of radiographs taken with γ rays may be less than of those taken with x rays since Ra capsules may be rented and are easily portable. D. B. DILL

An attempt to demonstrate the existence of short-range α-particles from radium C. G. H. HENNINGSON AND J. L. NICKERSON *Proc. Trans. Amer. Scian. Inst. Sci.* 17, 291 8 (1930) —An attempt was made to detect short range particles from Ra C by the Wilson chamber method but no group of particles of definite range could be detected above the general background. RACHIN BROWN

The rate of flow of heat of some radioactive minerals. A. DORANIALSKA *Bull. intern. acad. Polonaise* 1930A, 42 0 (in French) —See *C. A.* 24, 1280. J. W.

Application of an adiabatic microcalorimeter for measurements of rates of flow of heat of uranium, thorium and radioactive minerals. A. DORANIALSKA *Bull. intern. acad. Polonaise* 1929A, 487-96 (in French) —See *C. A.* 24, 1012. J. W.

Photographic effects of γ-rays. J. S. ROGERS *Proc. Phys. Soc. London* 43, Pt. 1, 59 17 (1931) —γ-Rays from Ra and Rn in equilibrium with their short lived products were collimated by Pb slabs and allowed to fall on Agfa duplitized film. Besides the direct beams, rays filtered through lead screens, varying in thickness from 0.19 to 4.03 cm., were used. The densities of the traces on the films were measured with a Moll microphotometer. Results are given for the variations of photographic density both with time and with intensity of radiation. The shape of the characteristic curves was found to be independent of filtering. Similar results were obtained for the Hurter-Dunfield curves, except when no half filter was used. The absorption coeff. of Pb was found to be 0.83 cm^{-1} for thicknesses from 1 to 7 cm. W. W. STRICKER

Ionium in dispersed masses of Tyuya-Muyun. Description of the method of preparation of a solution for its determination. I. D. KURBATOV, N. A. KARTHAVINA AND N. A. SAMOILO *Compt. rend. acad. sci. U. R. S. S.* 1930A, 69-74 —In certain ores of Tyuya-Muyun, Ra is shown to be present in excess of the amt. in radioactive equil. with U. It is assumed that this is due to the presence and transformation of Io. As there is no direct method for detn. of Io an indirect method was devised to solve the problem whether Io is in excess of the equil. ratio of Io:U. The sample was fused with Na₂CO₃ and a Ra salt. Ra and Io were sep'd as sulfates. The initial amt. of Ra and that resulting from transformation of Io were det'd with a Wulf electrometer. The results confirmed the assumption. Still another mineral deposit of In and Ra probably exists in Tyuya-Muyun. J. G. TOLPIN

Proportions of active elements in dispersed masses of Tyuya-Muyun. I. D. KURBATOV, *Compt. rend. acad. sci. U. R. S. S.* 1930A, 452-8 (1930) —Ra ore in Tyuya-Muyun consists of (a) active minerals contg. Ra, (b) minerals that contain no Ra but exert no influence upon the soly. of a in acids and (c) minerals contg. no Ra and lowering the soly. of Ra-contg. minerals. Only minerals of class a are discussed here. Most important of them is tyuya-muyunite, CaO·2UO₃·V₂O₅·xH₂O, which is easily sol. in dil. acids. Another Ra bearing mineral with entirely different chem. properties is brass-yellow radiobarite. From disturbed radioactive equil. it was concluded that other Ra-bearing minerals could be found in Tyuya-Muyun fields. For exptl. work active dispersed masses were chosen. Detn. of U and Ra showed an excess of Ra over the equil. amt. from 0.25×10^{-1} to 0.50×10^{-1} for each g. of U. The excess is explained as being due to sep'n. of Ra and Io from the ore in quantities in excess of the equil. amt. or to disturbance of equil. on account of migration of the active element with the sep'n. of a new mineral contg. Ra and Io formerly present in U mineral. In the second case the new mineral should be less sol. than tyuya-muyunite. Exptl. evidence tends to show the opposite. At least 3 minerals in which Ra is being conc'd. are present in Tyuya-Muyun; (1) tyuya-muyunite where Ra is stored ($U \rightarrow Io \rightarrow Ra$); (2) radiobarite ($Io \rightarrow Ra$), (3) dispersed Fe₂O₃ in primary ores ($Io \rightarrow Ra$). J. G. TOLPIN

Optics of x-ray reflection from crystals. V. Wide-angle diagrams. II. SzmMANN *Ann Physik* (5), 7, 633-49 (1930); cf. C. A. 25, 870 C. E. P. JEFFREYS

Studies in x-ray diffraction. I. The structure of amorphous carbon. P. KRISHNAMURTI *Indian J. Physics* 5, 473-86 (1930) —X ray powder diffraction patterns were measured for amorphous carbon prep'd from benzene, sugar, C_2Cl_4 and CCl_4 for active charcoal and for Acheson graphite. A strong scattering extending to an angle of 7° was observed for amorphous C which indicates that the C atoms are joined together in clusters of about 60 atoms of C per unit. II. Some colloidal solutions and liquid mixtures. *Ibid* 489-98 —Aq solns of starch, tannic acid and gum arabic show a strong scattering at small angles to the primary beams due to dissolved mols or micelles. The mol wt's calcd from the extent of the corona were 6200, 3131 and 2810, resp. The small angle scattering from $AcOH$ disappears immediately on diln indicating a breaking up of mol assocn. Mesitylene-cyclohexane, and ethyl phthalate benzene systems were also investigated. III. Aromatic hydrocarbons in the solid and liquid states. *Ibid* 543-58 —Acenaphthene, naphthalene, fluorene and phenanthrene show 2 rings in the liquid state while biphenyl, bibenzyl and stilbene show only one. An explanation for this is given based on the assumption of impedance to free rotation or vibration for compds having *o*- and *m*-substitution in the aromatic nucleus. The relation of the liquid diffraction patterns to crystal structure data is discussed. MALCOLM DOLE

X-ray diffraction in heated liquids and in solutions. V. I. VAIDYANATHAN *Indian J. Physics* 5, 501-21 (1930) —This work was done to test the applicability of the Einstein-Smoluchowski theory of optical scattering to small angles of scattering with x rays. Since, according to this theory, the ratio of the intensity of the scattered to the incident beam must increase at high temps., the x ray scattering of a no. of liquids was exam'd at high temps. With benzene the ratio is greater than the ratio of scattering in the optical region. The abs. values of the ratio for water and benzene at room temp. were calcd. theoretically and shown to agree well with the observed values. The origin of the inner and outer rings is considered. Since the inner halo intensity did not diminish with rise of temp. with assoc'd liquids like phenol the conclusion is that the inner ring origin is not due to assocn. The theory of scattering by dissolved mols is also discussed from a math standpoint. MALCOLM DOLE

Pari-absorption phenomena of x-rays. R. C. MAJUMDAR *Nature* 127, 92 (1931) —Partly absorbed lines were observed with C, N and Al as the absorbing substances and the incident radiations from $^{40}K_{\alpha 1}$, $^{40}K_{\beta 1}$ and $^{60}Ni_{K_{\alpha 1}}$. Cf. Ray, C. A. 24, 6008 J. P. JOHNSON

Influence of the form and polarity of molecules on the Röntgen spectrum of liquids. III. The appearance of two amorphous rings in substances whose molecules are probably disk-shaped. J. R. KATZ AND J. SELMAN *Z. Physik* 66, 834-57 (1930); cf. C. A. 22, 2881 —The Röntgen spectra of many fluid org. compds. have been observed and are found to consist of 2 "amorphous" rings sep'd by a zone of more or less diffuse light. Measurement of the rings gives data from which the structure of the mol. groups is inferred. The mols. are regarded as disk-shaped arranged in parallel layers very similar to a stack of coins, the stacks forming a space grating. This arrangement is disturbed by thermal excitation whereby the axes of the groups assume a random distribution and individual mols. may wander from one group to another. This is betrayed by a loss in distinctness of the rings in the Röntgen diagram. C. C. KIRKS

The intensity relations in the spectrum of a mixture of alkalis and the possibility of a quantitative spectrum analysis of these elements. HOWARD LUCAS *Z. anorg. allgem. Chem.* 195, 321-37 (1931) cf. C. A. 24, 5633 —The arc and spark spectra of salts of Na and K mixed in various proportions were photographed together with a source of white light of known intensity distribution to study the possibility of making quant. spectrochem. analyses. It was found that the relative intensities of the lines, as det'd. from microphotometer measurements, depended on the region of the arc under investigation. In the middle zone of the arc where the temp. and elec. field are lowest the spectrum of K is emitted with greater strength than that of Na. Detns. of the relative concns. of Na and K from the line intensities were made with an accuracy of approx. 20%. As a by product of the investigation it was found that the relative intensities of the principal and subordinate series are 1:100. C. C. KIRKS

The dispersion of internal energy between the quartet and triplet terms $3s^2P$, $3p^2P$, and $3d^2D$ of the spectra of carbon, nitrogen, oxygen and fluorine in different stages of ionization. J. GILLES *Compt. rend.* 192, 350-2 (1931) —From a study of the sepns. shown by the triplet and quartet *P* and *D* terms of C, N, O and F, coming from the electron configurations $1s^2 2s^2 2p^3$ and $1s^2 2s^2 2p^2 3p$, resp. It was found that (1) in passing from the 4th to the 6th column of the periodic table the triplet sepns. are very

nearly 80×2 and 80×4 wave no. units; (2) from the 5th to the 7th columns the quartet septns. are very nearly 80×3 and 80×5 units. C. C. KRESS

Investigations of the anomalous dispersion of excited gases. V. Negative dispersion in excited neon. H. KOPFERMANN AND R. LADENBURG. *Z. Physik* 65, 167-88(1930); cf. *C. A.* 22, 2514.—The neg. dispersions of the Ne lines resulting from the $s_1 - p_2$ combinations were measured by the method previously described (*C. A.* 22, 4364). With increase in strength of the current in the discharge tube the anomalous dispersions of all the lines terminating in the s_1 state increase in the same manner, but when the c. d. reaches 0.1 amp./sq. cm. the dispersions begin to decrease by amounts different for each line, depending on the atom density of each of the p_2 states. The no. of atoms in each of the p_2 states appears to reach a satn. value for a max. current of 0.7 amp., apparently because of the equil. between the atom densities of the various p_2 states and the exciting and neutralizing electrons. The sp. temp. for the s and p states derived from the measurements for max. current is approx. $23,000^\circ$. VI. Control experiments for demonstrating negative dispersion: absorption, anomalous dispersion, intensity distribution and intensity of different neon lines. R. LADENBURG AND S. LEVY. *Ibid.* 189-206.—Exptl. tests of the above interpretation of the anomalous dispersion of Ne lines is afforded by (1) the anomalous dispersion and absorption shown by lines resulting from combination of a p_2 state with a higher state, (2) the increase in intensity of the $s_1 - p_2$ lines with increase in strength of the exciting current. Measurements of the anomalous dispersion and absorptions of the line at 703.9 A. U. ($p_2 - s_1$) show that they increase markedly when the crit. current strength is reached and investigations of the intensity distribution within the red line at 630.4 A. U. by means of the Fabry-Perot interferometer show that it is caused solely by a Doppler effect. A formula is derived for the intensity of a spectrum line affected by absorption within a finite gaseous layer. C. C. KRESS

Two remarks on the dispersion of mercury vapor. R. LADENBURG AND G. WOLFSOHN. *Z. Physik* 65, 277-8(1930); cf. *C. A.* 24, 531.—(1) The new dispersion measurements in Hg vapor give a value of 1.14×10^{-10} sec. for the life of the 1P_1 state. (2) The frequency corresponding to the limit at 1190 A. U. of the principal series of Hg has assoc. with it a large f value (2.59). This is attributed to the first simultaneous jump of both valence electrons. C. C. KRESS

The variations in intensity distribution in a resonance spectrum. LEONID N. NIKOLSOV. *Z. Physik* 65, 75-83(1930), cf. *C. A.* 24, 5220.—The resonance spectra of Te and Se, excited by light from a Hg lamp, were observed at different pressures. The spectrograms, exposed also to the light of a W-filament lamp, were microphotometered to study the changes in intensity in the members of the resonance spectrum caused by the change in pressure. For Te intensity changes were noted which could not be attributed to absorption of the resonance rays by the vapor itself. For Se, in which the doublets of the resonance series were resolved by the spectrograph employed, the long wave length component of each doublet was strengthened relative to the short wave component with increase in pressure, the effect of absorption being eliminated. C. C. K.

Quantized rotation of the potassium atom. R. G. LOTARTE AND R. GRINFELD. *Univ. Nod. La Plata, Estud. Cienc.* No. 87, 103-8(1929), *Science Abstracts* 33A, 645, cf. *C. A.* 24, 2699.—Williamson has published curves which show the stimulation of K atoms by electron shocks. These curves show that it is probable that there is a quantized rotation of these atoms. The crit. potential, which corresponds to the leap from level $n = 0$ to level $n = 1$, must be 3.29 v., which corresponds to a $\Delta E = 27,441 \text{ cm}^{-1}$. On spectroscopic testing, 23 lines of the arc and spark spectra of K are found, obtainable by combination of a rotation-quantum leap with the electron leaps. Since the value $\Delta E = 27,441 \text{ cm}^{-1}$ is very large, it is not possible to find lines for which to subtract the rotation-quantum leaps from those of the electrons, for such lines would be in the deep infra-red, where measurements have not been made. The no. of combination lines within the observed region is, therefore, small. G. G.

Anomalous Zeeman effect. P. DRUMAUZ. *Bull. assoc. ing. Ec. Electroch. Mexico* 7, 101-8, 142-56, 164-72(1929), *Science Abstracts* 33A, 678.—This paper, in 3 serial portions, gives a summary of the theory of the anomalous Zeeman effect. The 1st portion deals with the classical theory, the 2nd with the classical quantum theory and the description of the effect in different multiplets and the final portion gives the quantum-mechanics treatment, with the use of wave mechanics and the matrix mechanics of Heisenberg and Dirac. G. G.

Elementary theory of the Zeeman effect. J. ZAHRAVÍČEK. *Z. Physik* 62, 694-5(1930).—The Lorentz equations for the motion of a bound electron in a magnetic

field are best solved by the usual transformation of 2 real equations into one imaginary equation

The Zeeman effect of quadrupole lines. E. SEGRI. *Z. Physik* 66, 827-9(1930) — The narrow doublet of K at 4042.17 and 4641.53 Å representing the transition $^2S - ^2D$ results from quadrupole radiation. The Zeeman effects of these lines have been observed in a field of 12,700 gauss and confirm not only the theory of Rubinowicz for the Zeeman effect of quadrupole radiation but also the character of these lines as forbidden $S - D$ combinations

The Stark effect of some helium lines in the visible part of the spectrum. KARL SJÖGREN. *Z. Physik* 66, 377-88(1930) — The canal ray method of Stark was used to observe the displacements of visible He lines produced by elec. fields ranging from 175 to 250 kV/cm. Details of the measurements are given for the lines $2S - 3P$, $2S - 3D$, $2P - 4S$, $2P - 4D$, $2P - 4F$, $2P - 4B$, $2p - 4s$, $2p - 4d$, $2p - 4f$ and $2p - 4p$

Lithium-like spectra of carbon, nitrogen and oxygen, C IV, N V and O VI. BEN EDLÉN AND A. LARSON. *Z. Physik* 64, 64-74(1930) — The spectra of C, N and O have been examined in the extreme ultra violet, the vacuum spark method being used. The series of Li like doublets was obtained. For the fundamental term $2^2S_{1/2}$, for the different elements, the following values were obtained: C IV, 520, 220; N V, 789, 591; O VI, 1,114, 206. These correspond with ionization potentials of 64.19, 97.425 and 137.482 v, resp. The term values for the complete series for Li I, Be II, B III, C IV, N V and O VI are compared

C IV lines in the visible and near ultra-violet regions of the spectrum. The term system of C IV. BENGT EDLÉN AND JOHN STENMÅN. *Z. Physik* 66, 328-38(1930) — With the terms of C IV previously described (preceding abstract) the lines which should appear in the near ultra violet and visible regions of the spectrum have been calculated. These lines and others belonging to C IV have actually been observed in the spectrum emitted by a condensed spark discharge between graphite electrodes in vacuum. These new wave length observations together with those in the extreme ultra-violet have permitted a recheck of all the terms up to those with principal quantum no. 6. The 2P terms follow Landé's formula for the doublet separ. very closely, giving a value 1.82 for the screening const. σ . C IV is now the best representative of a Li like spectrum

The nuclear moments of Li^+ and Li^0 . HERMANN SCHÜLER. *Z. Physik* 66, 431-3(1930), cf. C A 25, 249 — With a Perot-Fabry etalon and a Schüller lamp, the hyperfine structure of the Li II triplets, $2^2P_{3/2,1} - 2^2S_1$ at 5485 Å U_{vac} was photographed. Several discrepancies in the earlier measurements have been removed. The results do not agree with the calculated nuclear moments, $\mu = 1/2$ or $3/2$, but the latter is slightly favored. No splitting of the lines for Li^+ was detected. In case it exists, it cannot be more than $1/2$, as great as that for Li^0

The para-ortho splitting and the mean value of the S -terms of helium for higher quantum numbers. LONN A. HILLERÅAS. *Z. Physik* 66, 453-70(1930), cf. C A 25, 1156 — Math. By a perturbation method, the term values for the S states of para and ortho-He are calculated for higher quantum nos. A similar calcn. is made including the influence of the polarization of the atom core. The results are $E(n,0) = -R\lambda/(n - \sigma)^2$, where $\sigma = 0.176 \pm 0.054$, 0.22 ± 0.07 , 0.219 ± 0.079 for calcn. without and with polarization effect and exptl., resp.

The calculation of the spectra of ions. ENRICO FERMI. *Mem. accad. Italia. Classe sci. fis. mat. nat.* 1, No 2, 10 pp (1930) — F. first presents his general method of calcg. the statistical distribution of the electrons in an ion. The electrons surrounding the nucleus are treated as a cloud of electronic gas the d. of which is detd. statistically as a function of the distance from the nucleus. Next is given the calcn. of the elec. potential in the interior of an ion as a function of the distance from the nucleus. This makes possible the numerical calcn. of the energy levels, which procedure is exemplified by the calcg. of the S terms, or rather their Rydberg corrections, of the spectra of triply ionized atoms. Good agreement between calcd. and observed values is shown

The magnetic moments of atomic nuclei. ENRICO FERMI. *Mem. accad. Italia. Classe sci. fis. mat. nat.* 1, No 1, 12 pp (1930) — The hyperfine structure of spectra is due to the interaction between the motion of the electrons of the atom and the intrinsic magnetic moment of the nucleus. Methods are derived theoretically for calcg. this hyperfine structure of spectral lines. The magnetic moments of the nuclei of Na, Rb and Cs are deduced by applying these theoretical results to the experimentally detd. hyperfine structure of the spectra of these elements

Nuclear spin and the third law of thermodynamics. The entropy of iodine. W. F. GIAUQUE *J Am Chem Soc* 53, 507-14(1931), cf *C A* 25, 807—The value of $S_{298.15} = 27.9$ calories/degree per mole of solid I obtained by extrapolation of existing data represents the abs entropy less the nuclear spin entropy. For gaseous I the combined translational, rotational and vibrational entropies consists of the abs entropy less the nuclear spin entropy and has a value as obtained from band spectra, $S_{298.15} = 62.29$ cal/degree/mole. The high temp entropy effects due to nuclear spin in I are found to persist below 10^4 °K, which is not the case for H. The vapor pressure of solid I is given by the equation, $\log_e P(\text{atm}) = -(3512.3/T) - 2.013 \log_e T + 13.374$

A. LLOYD TAYLOR

Nuclear spin of nitrogen. W. R. VAN WIJCK. *Arch neerland sci. IIIA*, 13, 29-57 (1930)—In certain band spectra of mols in which the atoms are identical, alternate lines are very weak or even missing. This is ascribed to the symmetry of the mol and to resonance in the at nuclei. The theory of this method of investigating the nucleus, by photometric measurements of the intensities of the lines in the band spectrum, is outlined with special reference to the N_2 mol. Exptl details are given for work carried out on spectra emitted by the N_2^+ and the neutral N_2 mols. The ratio of the alternating intensities of the neg bands at 3914, 3884, 4278 and 4237 Å U was found to be 2:1. A Maxwell Boltzmann distribution among the rotational levels was found. The ratio for the nuclear spin was confirmed by other measurements, including work on the resonance line of Th at 3776 Å U.

W. W. STIFLER

The experimental assignment of the H_2 bands to the singlet and triplet systems. W. FINKELNBURG. *Z Physik* 66, 345-9(1930)—A spectrogram of H_2 , (obtained by Gehrecke and Lau) excited by streams of electrons of various velocities shows the same effect as that described for the singlet and triplet lines of Hg by Schaffernicht (*C A* 24, 5621). The lines which extend entirely across the spectrogram with undiminished intensity as the excitation energy increases are those belonging to the A, B and analogous bands of the singlet system. Those lines extending only part way across the spectrogram, fading out with increasing energy of the cathode rays, belong to the α, β, γ and δ bands of the triplet system.

C. C. KIESS

Optical dissociation of diatomic molecules in gases and vapors. I. II. III. G. H. VISSER. *Chem Weekblad* 27, 237-41, 380-4, 390-3(1930)—A survey of recent work.

B. C. A.

Analysis of the S_2 spectrum. R. ROMPE. *Z Physik* 65, 404-29(1930).—This analysis of the S_2 spectrum includes (1) the bands classified by Rosen (*C A* 22, 2717, 23, 564) which are shown to represent a $^1\Sigma - ^1\Pi$ transition, (2) the ultra violet bands described by Gilles (*C A* 23, 4406) which result from the term combinations $^1\Sigma - ^1\Sigma$ (or $^1\Sigma$) and $^1\Sigma - ^1\Pi$, (3) the resonance bands excited by the ultra violet rays of Hg. Five resonance series have been observed. These have been arranged into band schemes for which the corresponding anti-Stokes members have also been observed. The intensity distribution within the resonance series resembles that noted by Rosen for the fluorescence spectrum and deviates markedly from that required by Condon's theory. The effect of the addition of a foreign gas, N_2 or He, to the S_2 vapor is to bring out more or less completely the band systems associated with the resonance series.

C. C. KIESS

The emission bands of sulfur. R. K. ASCHUR. *Nature* 127, 93-4(1931)—Spectrograms of the emission bands of S were obtained from discharge tubes containing S vapor in the presence of high pressure A. New bands toward the red indicate that A helps in some way to prevent the predissociation of the S mols. A few new bands degraded toward the shorter wave lengths were also obtained in the extreme ultra violet beginning from about 3200 Å U.

W. F. JONESSTONE

Raman spectra of some triatomic molecules. S. BHADAVANTAM. *Nature* 126, 995 (1930)—Photographs of the Raman spectrum of SO_2 gas show a line with a wave no shift of 1154, which differs somewhat from the strong line having the shift 1146 which was previously found in the Raman spectrum of liquid SO_2 . The 2 fainter lines found in the spectrum of the liquid do not appear in that of the gas. Photographs of the spectrum of liquid HCN show a triplet with shifts of 2076, 2097.2 and 2122, the middle component being much stronger than the others. The results of Krishnamurti on CS_2 were confirmed and 3 new, very feeble bands at 4438, 4605 and 4680 Å U. were found in the spectrum of CS_2 excited by the Hg line at 4358.3 Å U.

W. W. STIFLER

Raman effect and molecular structure. III. The basic frequencies of molecules of the type XY_4 . Mixed halides. B. TRUMPY. *Z Physik* 66, 790-806(1930), cf *C A* 25, 250—The Raman spectra of CCl_4 , $SiCl_4$, $TiCl_4$, $SnCl_4$, $SnBr_4$, and of mixts of $SnCl_4$ and $SnBr_4$, in the ratios 1:1, 3:1 and 1:3, were observed. The measurements show that the tetrachlorides and also the tetrabromide all exhibit 4 basic frequencies, those of

SnBr_2 being displaced to longer wave lengths compared with those of SnCl_4 . These observed frequencies are in good agreement with those calculated from a theoretical consideration of a mol having its outer atoms at the vertices of a tetrahedron. Two of the frequencies represent active oscillations and 2 inactive oscillations of the mol. The infra red absorption spectra of these compounds, so far as known, correspond to the active oscillations and to combinations of them. The Raman spectra of the mixtures show new lines in addition to those due to the constituent mols. The lines of the original mols are weak compared to those of the new mols, which are probably of the form SnCl_3Br , SnCl_2Br_2 , and SnClBr_3 . C. C. KISS

Apparatus for the demonstration of the Raman effect in liquids. N. B. REYNOLDS AND F. BENFORD. *Rev. Sci. Instruments* 1, 413-6(1930).—A hot-cathode He lamp (of cylindrical shape) is used in conjunction with a Ni oxide glass filter (cf. Wood, *C. A.* 23, 5419). The lamp is surrounded by a reflector of elliptical cross section, the cross section of the lamp being concentric with one of the foci of the ellipse. Concentric with the other focus is placed the tube containing the liquid to be examined. A hot-cathode Hg lamp has been similarly used. B. C. A.

The Raman effect and the polymerization of water at various temperatures. O. SPECCHIA. *Nuovo cimento* [N. S.] VII, 10, 389-91(1930).—At room temp. the Raman spectrum of water excited by means of a Hg arc is over the range 4108 to 5790 Å. U. gives 3 bands of greatest intensity at 4176, 4693 and 5158 Å. U., and 2 much weaker bands at 4258 and 4480 Å. U. The bands at 4176 and 4693 Å. U. are composed of 3 components each. The effect of increase of temp. in modifying the band at 4176 Å. U. was particularly studied. Measurements were made between 17° and 91°. A marked diminution of intensity of the 3 components of this band was noted with increasing temp., as well as a displacement toward greater wave length. L. T. FAIRHALL

The photochemical union of hydrogen and chlorine. I. The effect of light intensity. ARTHUR J. ALLMAND AND EDWARD BEESLEY. *J. Chem. Soc.* 1930, 2003-708.—An extensive study to det. the relation between the reaction velocity and the intensity of light in the photochem. combination of H_2 and Cl_2 . Thus, 405 mμ and 436 mμ monochromatic light over measured intensity ratios of about 180:1 and 440:1 resp., plane-polarized complex light over a measured intensity ratio of 4:1, and intermittent monochromatic light of wave lengths 313 + 303 mμ, 365 mμ, 405 mμ and 436 mμ are used. The light intensities are measured with a thermopile. They are varied in the expts. with the 405 mμ and the 436 mμ monochromatic light by 4 wedges of neutral tint and continuously varying density, with the polarized light by crossed Nicol prisms, and in the expts. using the intermittent monochromatic light by a rotating disk with 2 adjustable open sectors in opposite quadrants. The results show that the reaction velocity is proportional to the intensity of the incident light. The results of Baly and Barker (*C. A.* 24, 2385) are explained. J. BALOZIAN

The photochemical union of hydrogen and chlorine. II. The effect of wave length. Measurements with filtered light. ARTHUR J. ALLMAND AND EDWARD BEESLEY. *J. Chem. Soc.* 1930, 2709-21.—The dependency of the quantum efficiency (γ) on the wave length (260-546 mμ), the relative temp. coeffs. for the lines 436 mμ, 405 mμ, 365 mμ and 313 + 303 mμ (comparative velocity measurements being made at 19.7° and 25.0°), and the simultaneous effect of 2 monochromatic rays are investigated in the photochem. reaction between H_2 and Cl_2 . The app. and methods are similar to those used in expts. of the preceding abstract, extra light filters (260 mμ and 546 mμ) being also used. With decreasing wave length, γ at first increases to a max. and then decreases rapidly in the ultra violet. The expts. show that between 313 mμ-436 mμ the temp. coeff. of γ increases with increasing wave length. In every case the simultaneous effect of 2 monochromatic rays is equal to the sum of their sep. effects. The threshold wave length for this reaction is found to be 546 mμ. J. BALOZIAN

The photo-reaction of hydrogen and iodine monochloride. D. P. MELLOR AND T. IREDALE. *Nature* 127, 93(1931), cf. Rollefson and Lindquist, *C. A.* 24, 4224.— H_2 and ICl react rapidly in strong light if the H_2 pressure is large compared with the pressure of ICl . This indicates that conditions favor the greater probability of the reaction, $\text{H}_2 + \text{Cl} \rightarrow \text{HCl} + \text{H}$, and that excited Cl atoms may not be necessary for it to take place. In thin glass bulbs the reaction between H_2 and ICl is very slow in artificial light, more rapid in diffuse daylight, and very rapid in direct sunlight, I_2 and HCl being the main products. H. F. JOHNSTON

Photo-reaction between hydrogen and iodine monochloride. G. K. ROLLEFSON AND F. E. LINDQUIST. *J. Am. Chem. Soc.* 53, 1184-5(1931).—R. and L. point out that the results of Mellor and Iredale (preceding abstract) are not inconsistent with their

previous conclusions (C. A. 24, 4224), since their exptl. conditions were considerably different

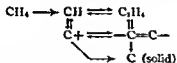
Photochemical interaction of ethylene and ammonia. HUGH S. TAYLOR AND HARRY J. EMLEUS. *J. Am. Chem. Soc.* 53, 562-74(1931).—The photochem. decomn. of NH_3 in the presence of C_2H_4 and C_2H_2 - H_2 mixts induces a polymerization of C_2H_4 at room temp. to a liquid product. A formation of satd hydrocarbons is subordinate to this process. The velocity of the reaction is independent of the C_2H_2 pressure over a wide range, and of NH_3 pressure when absorption is complete. Conclusions are reached as to the efficiency of methods for producing satd hydrocarbons from C_2H_4 and at H_2 .

Photochemical decomposition of hydrogen peroxide in aqueous solution in presence of sodium nitroprusside. I. M. QURESHI. *J. Phys. Chem.* 35, 656-8(1931).—An aq. soln. of H_2O_2 to which a few drops of Na nitroprusside soln. are added is sensitive to the visible light, the decomn. of the H_2O_2 continuing after the illumination ceases. It is probable that the illumination produces some colloidal Prussian blue which acts as a catalyst for the dark reaction. The course of the reaction was not followed quantitatively but this will be done later.

The transfer of energy between atoms at collision. OSCAR KNEPLER RICE. *Proc. Nat. Acad. Sci.* 17, 34-9(1931).—Kallmann and London have considered the question of energy exchange between atoms and mols. at distances greater than those comparable with the at. or mol. radii. They neglected the relative translational energy of colliding atoms or mols. A modification of Born's method for solving collision problems is outlined which takes into account the mutual kinetic energy of the particles and meets the objection of K. and L. (C. A. 24, 4696) to Frenkel's suggestion (C. A. 24, 2032).

The separation of two types of iodine molecule and the photochemical reaction of gaseous iodine with hexene. R. M. BADGER AND J. W. URMSTON. *Proc. Nat. Acad. Sci.* 16, 803-11(1930).—Since one type of I mol. absorbs the Hg line 5461 and the other does not, an attempt at sepn. has been made by causing the activated type to react with hexene vapor. Indications, which are not conclusive, point to a partial sepn.

The mechanism of the action of electrical discharge upon methane. KURT PETERS AND O. H. WAGNER. *Brennstoff-Chem.* 12, 67-8(1931), cf. C. A. 23, 3861, 4326, 24, 2935, 5715, 25, 399, 645.—Spectrum analysis of the arc in the discharge tube before the color change observed upon increasing the elec. charge shows the bands of the $-\text{CH}$, N and $-\text{CN}$ groups. This is taken to show the progression in radicals from $\text{CH}_3 \rightarrow \text{CH}_2 \rightarrow \text{CH}$. Polymerized satd hydrocarbons are thus formed at low pressures and low charges while C_2H_4 and C_2H_2 are formed at high pressures. Following the change from pale blue to yellow, bands of the Balmer series, the Swann spectrum with exception of group I, and the $-\text{C}=\text{C}-$ and $\text{C}+$ groups were observed. Solid hydrocarbons result from the reaction of $-\text{C}=\text{C}-$ groups with reactive H. The mechanism is indicated as



The relation between elec. consumption, % in the final gas and the charge is shown graphically for C_2H_2 and C_2H_4 .

The ammonia discharge tube. G. I. LAVIN AND J. R. BATES. *Proc. Nat. Acad. Sci.* 16, 804-8(1930), cf. C. A. 24, 3441.—The active gas reacts with ethylene to give cyanides and with O to give nitrides. Solid surfaces show sp. action. Cu and Fe_2O_3 extinguish the glow immediately. Sn, Ni, Fe, Pt, show an induction period. Mo and W produce little effect.

Disadvantages of the quartz lamp in fluorescence analysis. J. PLOTNIKOV. *Chem. Zig* 54, 582(1930).—The low light intensity, fragile nature, and monochromatic emission are mentioned, only those substances fluoresce under the lamp which can absorb radiation of wave length 366 mμ.

Transmission changes in ultra-violet glasses during high-temperature exposure to light. C. C. NITCHIE AND F. C. SCHMUTZ. *Science* 71, 590(1930).—When kept in contact with the hot tube of the Hg arc lamp (about 450°), certain glasses showed a marked increase in short-wave transmission.

The absorption of ultra-violet rays by liquids transparent to light. J. CLUZET

AND T. KOFMAN *Compt rend soc biol* 103, 783-5(1930).—A Cd photoelec. cell was used to measure the ultra violet light (2000-2300 Å U) absorbed by several liquids transparent to visible light B. C. DARNSTETTER

Quantitative investigations on the absorption spectra of organic dyes. J. AIZKARA, *Isla Univ Latvianis Kim Fakultat Serija I* (in German) 362-8 (in Lettish) 279-301.—Absorption spectra of 79 org dyes were measured with a Koenig-Martens spectrophotometer between 7200 and 4100 Å U. Values of the mol. absorption const. are given for each. The advantages of the quant. method are discussed. The Lambert-Beer law is valid for weak dye solns in presence of small amt. of colorless inorg. compds. The results are discussed with regard to the relation between spectra and structure of the dye. The advantages of the color triangle are also discussed. G. M. MURPHY

Effect of impurities on the phosphorescence of calcium sulfide. D. N. GOVIL AND N. SINGH *J chim phys* 27, 433-51(1930).—The effect of adding up to 0.1% of eosin, fluorescein, quinine hydrochloride, methyl blue, $PbCl_2$, $Pb(NO_3)_2$, $CuCl_2$, $CuSO_4$, $Cd(NO_3)_2$, $CdSO_4$, $Ba(NO_3)_2$, $BaCl_2$, $Sr(NO_3)_2$, and $SrCl_2$ to CaS on the phosphorescence of the CaS was measured. The first 6 named gave a max. effect in concns. of about 0.003%, the highest value being obtained with quinine hydrochloride. Addns. of the other mineral salts gave minima in small concns., followed by maxima which were usually lower than the initial value. The intensity of the radiation decreased with age. The luminosity of CaS is explained by the equations: (1) $2CaS + 2H_2O \rightleftharpoons Ca(SH)_2 + Ca(OH)_2$, (2) $CO_2 + Ca(OH)_2 \rightarrow CaCO_3 + H_2O$, (3) $H_2O + CO_2 + Ca(SH)_2 \rightarrow CaCO_3 + 2H_2S$. H_2S was detected by odor and by test with lead acetate.

GERALD M. PETTY

Thermoluminescence in glasses which contain two activators. BYRON E. CONY AND WM. D. HASKINS *J Am Chem Soc* 52, 5144-54(1930).—The thermoluminescence of a zinc borate glass caused by Ce and Mg alone and mixed in concns. up to 1% of each element was measured from 2000 to 3350 Å U. The effect of mixts. was not additive. Intensity measurements of the thermoluminescence indicate that the energy for thermoluminescence is drawn from near the edge of the absorption bands. Relative thermoluminescent intensity varies with the source of excitation. The data are presented graphically. GERALD M. PETTY

Scattering of light by dielectrics of small particle size. G. F. A. STUZZ *J Franklin Inst* 210, 6, 85(1930).—The diffusion of light by a dielec. material Zn oxide, was investigated in the range of particle sizes where, with increasing particle size, Rayleigh scattering rapidly diminishes and reflection and refraction effects increase in importance. The turbidity max. for Zn oxide dispersed in water, detd. for 3 wave lengths of light occurs at 0.25 μ for the wave length 5500 Å U. and at smaller particle sizes for shorter wave lengths. The particle size of max. turbidity was found to be proportional to the wave length of the light in the range 4900-6300 Å U. An app. for detg. the angular distribution of the light scattered by a suspension of fine particles is described. The total observed intensity has not the same distribution as calcd. by Blumer (cf *Z Physik* 38, 304(1925)), the observed depolarization is in fair agreement with theory (cf Lange *L* 1, 22, 253-4). B. C. A.

The acceleration produced by light of the flocculation of colloidal solutions in fluorescent media. ALBERTIN BOUTARIC AND JEAN BOUCHARD *Compt rend* 192, 457(1931).—The effect of visible and of ultra violet light on the rate of flocculation of As_2S_3 and other sols by electrolytes in presence of fluorescein, eosin and erythrosin was studied. In all cases illumination decreased the time required for flocculation, the effect being greater in ultra violet than in visible light freed from infra red. K_2SO_4 and H_2SO_4 , which inhibit the fluorescence of the dyes also suppress the effect of light. Light has no effect in the absence of fluorescent compds., and the dyes do not affect the sols in the dark. K. V. THIMANN

Increased bactericidal effect of inorganic compounds in the presence of x-rays. ROBERT J. NOBIS *Bull Banc Sci Research* 3, 21-36(1931).—When bacteria are irradiated in solns. contg. $NaBr$, $NaCl$, NaI , $BaCl_2$, $HgCl_2$, $LiO_2(NO_2)_2$, CsI or $Th(NO_3)_3$, the resulting lethal action is very much greater than that obtained by exposing the bacteria successively to x rays and to the salts previously irradiated with x rays, and is a true synergistic action. This action has previously been ascribed to ultra violet fluorescence or to emission of photoelectrons. Expts. are described which indicate that the synergistic action of the above salts is not due primarily to these factors. The expts. indicate that the synergistic action is independent of permanent chem. decompn. For Th salts, it has been shown that at the crit. voltage for the ejection of the K electrons the resulting action is greatest, while theoretically the speed of the emitted K electrons is at a min. at this voltage, and that at a higher voltage, at

which K electrons are emitted at much greater speed, the action is much less. A tentative explanation is based upon the theory that the toxicity of the ion is related to its ionic potential or state of ionization. J. A. KENNEDY

Röntgenographic investigation of the system Cd-Mg (DEHLINGER) 2. An attempt to separate the isotopes of Br (HARTZKE, STRIEBEL) 2. A series of new tetramminecobaltic complexes (SARKAR, DAS-GUPTA) 6. Photoelectric cells in science and technics (LANGE) 4. Isotopes of K their association with plant life (DRECE) 11D. Ultra violet absorption spectra of the quinoline group (HICKS) 10. Isotopes and living organisms (VERNADSKII) 11D

Ergebnisse der exakten Naturwissenschaften. Band VIII. J. ESTERMANN Elektrische Dipolmomente von Molekülen. H. SACK Dipolmoment und Molekülstruktur. Berlin J. Springer Reviewed in *Z. physik. Chem.* 152, 318 (1931) HUMPHRIS, F. HOWARD Ultra Violet and Other Rays. London: Macmillan 123 pp

Electron-emitting compositions. ANTON LEDERER. *Austrian* 120,212, Sept 15, 1929. An electron-emitting compn. comprises more than 2% of Th incorporated into a metallic carrier consisting mainly of Os, with or without difficultly fusible metals or alloys such as W. The compn. may be coated on a wire or the like of W, or a paste contg. ThO₂, Os and a binder may be extruded to form a filament, which is then heated to a temp. above the m. p. of Th, suitably in an atm. of CO. Addnl. details are given.

Ionizing gases. ERIC DUTY. *Fr.* 695,016, Aug 5, 1929. Gaseous currents are ionized by bringing about chem. reactions at ordinary temps. in an elec. field produced by an alternating or pulsatory source of current the potential of which is not below 1500 v., so that the shocks of the ionized mols. between themselves liberate a sufficient amt. of energy to produce the reaction. The process may be used for the production of NO from N and O.

4—ELECTROCHEMISTRY

COLIN G. FINK

The direct-current arc furnace. SIEGMUND SCHIEY. *Grossere-Ztg.* 26, 583-7 (1929).—The elec. accessories of an elec. steel furnace are described. C. L. W.

The high-frequency furnace in the steel mill. N. BROGLIO. *A. E. G. Mitt.* 1931, 37-42.—The "coreless" high frequency furnace in which the charge (contained in the crucible) forms the core is described; it is used in an improved method of producing C and alloy steels. The layout of the instrument panel and elec. connections is given. A 300 kg. charge of cold high-speed steel was melted and poured in 82 min. with a requirement of 720 kw. hrs. per ton. C. J. B.

Use of electric furnace for the preparation of nonferrous alloys. K. OKIMOTO. *J. Fuel Soc. Japan* 10, 48-58 (1931).—The use of an elec. furnace in place of a coke-crucible furnace in prepn. of Cu alloys is recommended. F. I. NAKAMURA

Photoelectric cells in science and technics. B. LANGE. *Naturwissenschaften* 19, 103-7, 128-32 (1931), cf. *C. A.* 25, 1435.—A review with many references.

Photoelectric control for soaking-pit covers. R. M. BOYLE. *Elec. World* 97, 409 (1931).—A Middle West plant has a soaking-pit control which makes use of a light-sensitive relay as the initiating source. The moving crane carries only a foot-operated push button and a light box mounted on the end of the crane for each soaking pit spanned by the crane. The layout consists of 2 soaking pit cranes operating on a common runway and each spanning 2 pits. Preparatory to opening a cover or during the operation the crane may be shifted 5 ft. on either side of the center line and the operating lamp-box will still be opposite the light relay for that cover. To open the cover a single short flash of the lamps in the box is all that is necessary. The circuit is self-maintaining. A simple transfer circuit permits the light ray to be alternately effective to open or close the pit cover. The light-sensitive cells respond only to comed light above the room illumination, leaving no danger of false operation. W. H. BOYNTON

Temperature compensation in electrolytes. E. BLAMBERG AND K. MÜLLER. *Arch. Elektrotech.* 23, 435-40 (1930); *Science Abstracts* 33B, 411.—The fact that the cond. of an electrolytic cell varies rapidly with temp. as well as with concn. may be a disadvantage in some processes of deposition. The authors show how to proportion a

combined sheet waterproof resistance and a series Cu resistance to compensate substantially for the effect of temp. G. G.

Electrolytic cells for chlorine and caustic soda. C. L. MANTELL. *Chem. Met. Eng.* 38, 88-90 (1931).—Ten electrolytic cells for Cl—the Castner, Sorensen, Townsend, Hargreaves-Bird, Nelson, Allen-Moore, Bock-McRae, Gibbs, Wheeler, Vorce—are briefly described as to their constructional details, and compared. A table of operating data on the various cells has been tabulated from information furnished by users of the equipment. Yields of from 0.62 to 0.76 lb. Cl and 0.69 to 0.85 lb. NaOH per kw. hr are in the range of com. practice. C. L. MANTELL.

Electrolytic iron. R. DUPUIS. *Métallurgie* 26, 1829 (1933).—A lecture reviewing the qualities of electrolytic Fe and the technical methods of its production. L. P.

A note on the preparation of silver-free copper. BART PARK. *J. Am. Chem. Soc.* 52, 3165-6 (1933).—The last spectrographic traces of Ag were removed from Cu by electrolyzing a soln. of 4.0 g. CuSO_4 made up to 2 L. with 5% HNO_3 between Pt electrodes with 0.05 amp. at 7 v. The Ag in the 5 fractions of deposits removed at intervals of 1-24 hrs. was estd. by Nitsch's spectrographic method (*C. A.* 23, 1887). J. D. S.

Electrolytic gold plating with high current densities. H. FAWCETT AND R. WERNER. *Z. Elektrochem.* 36, 872-80 (1933).—The compn. of the bath recommended for Au plating with high c. d. is 200 g. $\text{K}_2\text{Fe}(\text{CN})_6$, 50 g. HAuCl_4 , 50 g. anhyd. Na_2CO_3 and 1 l. of H_2O . The ppt. of basic Fe carbonate is filtered from the warm soln. and the clear filtrate serves as the electrolyte. The bath is used at 70° with rapid stirring. Gold anodes are used with the surface equal to or > that of the cathode. The cathodic c. d. is maintained at 4-6 amp./sq. dm., which corresponds to an e. m. f. of 1.2-1.5 v. when the electrodes are 4-6 cm. apart. A smooth Au plate is produced which does not require burnishing. At higher c. d. the plate is darkened. With the recommended conditions the cathodic current efficiency is equiv. to 150-100%, and the anodic efficiency to 115-180%. A further study was made of the effect of varying the concn. of each of the components of the bath on the c. d.-anode potential curves. H. F. JONVREUX.

Electroplating wires and strips at high current densities. A preliminary communication. D. I. STEPANOV, B. N. KABANOV AND N. T. KUDRYAVTSEV. *Tekhnika Met.* 1930, 1151-8.—Cold Zn plating has many advantages over the hot dip process, such as greater uniformity of the coating, steadiness of the bath, economy of Zn and easier control of the bath. The proposed methods of increasing the c. d. to hasten the process have serious drawbacks. The authors made about 300 expts. to find conditions essential for good results with high c. ds. Circulation of the bath permits higher c. ds. (200 to 400 amp./sq. dm.) than those used before. Different concns. of ZnSO_4 , H_2SO_4 , H_2RO , making up the bath were also studied, as well as the changes in the bath during the process and the suitable indicators. At 50 amp./sq. dm. it is necessary to add to a bath of 1000 l. 50 l. water, 1 kg. H_2SO_4 , and 10 g. boric acid per hr. At 50 amp./sq. dm. the best temp. is 45°. At higher c. ds. the temp. should be higher. The problem of cooling the electrolyte which is heated by the current is solved by the application of air stirring; a formula is developed by which the equilibrium temp. suitable for the process can be calcd. Tables show the change of voltage with temp. and with the distance between the electrodes and a set of optimum conditions at c. ds. 50, 100 and 200 amp./sq. dm. J. G. TOLPIN.

Electrodeposition of lead-thallium alloys. C. G. FINK AND C. K. CONARD, JR. *Metal Ind. (N. Y.)* 29, 116-7 (1931).—See *C. A.* 24, 4989. G. G.

Polarographic studies with the dropping-mercury cathode. XVI. Electroreduction of acetaldehyde. I. SUMMER. *Canadian Chemical Chem. Comm.* 2, 699-71 (1930), cf. *C. A.* 24, 5579.—The detection and estn. of aldehydes occurring in certain alc. beverages by their electroreduction potentials, with the dropping Hg cathode, have been previously described (Shikata and Tachi, *C. A.* 21, 1917; Shikata and Shoji, *C. A.* 22, 2895). The aim of the present work was the investigation of the mode of reduction of AcH in its pure solns., and its detn. in fermentation products. The reduction of the AcH was studied in solns. of Li salts or LiOH, since Li has the most neg. deposition potential of all cations at the dropping Hg cathode. AcH was found to be reduced at a cathode potential 0.50 v. more pos. than that at which Li is deposited from solns. of equal concn. as the AcH solns. Thus the reduction potential of AcH in a 1 M soln. is -1.60 v. The value of the reduction potential was uninfluenced by changes from a neutral to a weakly acid, or to a strongly alk., soln. Increasing acidity of the soln. hindered the electroreduction. A change in concn. c. of AcH produced a shift of the reduction potential, v , according to the formula $v = 0.060 \log c$. Math. analysis of the bend of the current voltage curve showed that $v = -0.070 \log i$ (i = current), i. e., $v = -(RT/F) \log i$. The value of the diffusion current in 0.001 M AcH solns. was

the same as those given by 0.001 M solns of Ba or Sr chlorides; it is, therefore, concluded that the electroreduction of 1 g mol of Aell requires the same amt. of electricity as is required to deposit 1 g equiv of Ba, i. e., 1 Faraday. The probable reduction product is stated to be dimethylglycol. With this method it was found possible to est. the presence of 1 part Aell in 2 million to an accuracy of 10%. The results obtained were applied to the estn of Aell in Foll samples and in various vinegars. In fermented vinegars 0.001-0.002% Aell were estd. polarographically, while in synthetic or pyroligne vinegars no traces of Aell were found.

E. B. SANGAR.

Anodic oxidation of acetic acid in sulfuric acid solution. RUDOLPH SCHREINER. *Z. Elektrochem.* 36, 953-63 (1930).—In order to study the mechanism of the anodic oxidation of AcOH, solns. of oxalic, glyoxylic, glycolic, acetic and formic acids, HCHO and MeOH were electrolyzed, with a diaphragm, in 2 N H₂SO₄ at const. e d. Oxidation potentials were detd., and the gaseous products were analyzed. The electrodes were Pt foil. In the AcOH series the main reactions are:

$\text{CH}_3\text{COOH} \xrightarrow{+O} \text{H}_2\text{COH-COOH} \xrightarrow{+O} \text{HC(OH)}_2\text{COOH} \xrightarrow{-2\text{H}} (\text{COOH})_2 \xrightarrow{-2\text{H}} 2\text{CO}_2$, and the side reactions: $\text{H}_2\text{COH-COOH} \xrightarrow{-2\text{H}} \text{H}_2\text{CO} + \text{CO}_2$ and $\text{HC(OH)}_2\text{COOH} \xrightarrow{-2\text{H}} \text{HCOOH} + \text{CO}_2$. In the MeOH series, the main reactions are: $\text{CH}_3\text{OH} \xrightarrow{-2\text{H}} \text{H}_2\text{CO} \xrightarrow{+O} \text{HCOOH} \xrightarrow{-2\text{H}} \text{CO}_2$.

The data are interpreted as suggesting the following general rules for anodic oxidation of org. compds: (a) an increase in potential favors addn of O over the removal of 2H; (b) in general, the removal of 2H atoms from the same C is easier than the addn of O; (c) the difficulty of H removal increases in the order —OH, —CH, —COOH; (d) the simultaneous removal of 2H atoms attached to different C atoms (in the present case 1H was always a member of —COOH) is possible only when the H atoms are at the end of a 6-membered chain.

H. F. JONASTON.

Electrodeposition of chromium. B. RASSOW AND L. WOLF. *Chem.-Ztg.* 55, 73-6 (1931).—An investigation of the com. prepn. "Chromprotelt," used as a protecting layer to prevent formation of chrome mists, is reported. Cf. C. A. 24, 4993. C. L. W.

The acute effects of chromic acid mists. L. SCHWARTZ AND F. SEIDKE. *Zentr. Gewerbehyg. Unfallverhütung* 17, 232-4 (1930).—Mists given off by Cr plating baths are the cause of ulcerations and other skin diseases among the employees. Of 233 persons examd., 42.6% had dermatitis, ulcers or scars and 52% showed nasal changes; of these 37 had perforation of the septum, and 79 showed ulceration or devitalization of the mucous membrane. Ulcerations of the mucous membrane were seen after 2 weeks, and perforation of the nasal septum from 6 months to 4 yrs. The authors recommend exhaust ventilation or covering the surface with petroleum or some of the less dangerous homologs of C₁₂H₁₄. Objections to these methods are liability to fire hazard and C₁₂H₁₄ poisoning.

GEORGE R. GREENBANK.

Modern metal cleaning. LESLIE WRIGHT AND F. TAYLOR. *J. Electroplaters' and Depositors' Tech. Soc.* 6, 71-90 (1931).—This paper is a survey of work which has been done by various experimenters upon the nature of soap solns and the mechanism by which such solns remove dirt. Various types of soaps, their alky., phys. properties and cleansing properties are reviewed. The evaluation of the alky. and of the detergent action of soaps is explained. The application of detergent materials to modern metal-cleaning practice is discussed. The authors are of the opinion that no one chem. cleaner is yet available which will satisfy all plating-shop requirements, so that scouring and mopping and subsequent acid dips to remove tarnish are likely to remain in general use. However, with large batches of material, which have passed through similar shop processes, a study of the grease and dirt accumulated during these processes makes it possible to devise a cleaning soln. best suited to the particular foreign materials to be removed.

E. B. SANGAR.

High-voltage surge testing. F. D. FIELDER. *Elec. J.* 28, 181-4 (1931).—The cathode ray oscillograph operation is outlined—the importance of the timing systems being emphasized. The 3-electrode gaps permit the operation of the Dufour type oscillograph for the cathode voltage within the one microsecond. The operation is simple and consistent when care is used in the circuit design and adjustment. In synchronizing circuits, the most complex system is the combination of the Dufour type oscillograph and single-sweep timing. The essential parts of the system are shown in a fig.

W. H. BOYNTON.

Developments in hydrogen-cooled condensers. C. J. FRICHEDER. *Elec. J.*

28, 165-9(1931) — Success of the H_2 -cooled condenser is largely due to the development of many devices for making the machine safe. Features of the machine include (1) it has been made explosion proof, (2) an inert gas is used for scavenging when changing from air to H_2 or the reverse; (3) the internal pressure is automatically maintained above atm. pressure to prevent air entering; (4) various alarm signals are provided to inform the operator that the internal pressure or % of H_2 is low, or that any part of the machine or auxiliaries is not performing properly; (5) suitable indicating devices are provided for the gas, water, bearings, etc., in addition to the usual elec. instruments. Brief descriptions of these devices are given. Several illustrations and a good bibliography are included.

W. H. ROYTOR

Some comments on the use of "getters." GEORGE D. O'NEILL. *Electronics* 2, 510-11, 522(1931) — The application of the following materials as "getters" or "clean-up" agents in various types of radio tubes is considered. P, Mg, Al, Misch metal, Ba, Ca, Li, active charcoal and cerium.

M. McMANIS

Sensitizing the photoelectric cell. RICHARD FLEISCHER. *Electronics* 2, 518-9, 530(1931) — Four requirements of photoelectric cells are instantaneous response, linear relation between light intensity and resulting electronic current, sensitivity in the visible spectrum, and constancy. Data are given showing that to increase the sensitivity of the K layer in the photoelectric cell, the K in vapor form must be brought into contact with the gas (H_2 or inert gas) present in the tube. The K is vaporized by ionic bombardment.

M. McMANIS

The cleaning of blast-furnace gas by electric precipitation; the Siemens-Schuckert system at the Falva Works of the Bismarckhütte. HERMANN BASSER. *Stahl u. Eisen* 49, 1153-4(1929) — By elec. pptn. in 5 filter units, the suspended matter is sep'd. from 14,000 cu. m. of blast furnace (cu. m. at 0°, 760 mm.) gas at a flow of 3.5 m./sec. The cleaned gas has a solid content of 0.01 to 0.02 g./cu. m. The d. c. at 55,000 v. is obtained by means of a transformer from an a. c. at 500 v. The moisture content of the gas is very important, if it is too low, the particles are not charged properly. The required moisture content of the gas depends also on the compn. of the gas, a high Zn and Pb content requires a higher moisture content. The gas temp. must be kept above the dew point, in order to prevent clogging of the equipment. There is a pre-cooler installed before the pptg. filters, where the gas is cooled by injection of water, followed by heating of the gas attained through combustion of part of the gas. The cost of the elec. pptn. is 1/3 of that of the wet mech. pptn. at the Falva Works.

J. A. SELLARD

Electro-melted cement as a new building material in the chemical industry (Reviews) 20. The mechanism of the action of electrical discharge upon CH_4 (PATTERS, WAGNER) 3. The hardness testing of electrodeposits (O'NEILL) 9. Electro-furnace production of high heat-duty refractories (SCHROEDER) 19. Analysis of petroleum and its distillates for reducible substances and adsorbable matter by means of the polarographic method with the dropping-Hg cathode (GOSMAN, HAYROVSKY) 22. Evaporation of electrolytic caustic soda (LEE) 18. Values from ores by leaching and electrodeposition (Brit. pat. 336,584) 9. Magnetic testing system for electric coils (U. S. pat. 1,792,249) 9. Electrodeposition of rubber (Brit. pat. 337,639) 30. Alloys for sealing to glass such as in vacuum tube and incandescent lamp manufacture (Brit. pat. 337,680) 9. Recovery of Zn from ferrite compounds in the electrolytic Zn process (OLSHANSKY, NICONOFF) 9.

MARTI, OTTO MAR K., AND WINOGRAD, HAROLD. Mercury Arc Power Rectifiers. New York: McGraw-Hill Book Co. 473 pp. \$6. Reviewed in *J. Western Soc. Eng.* 36, 63(1931)

Gas-exit for electric batteries. FRITCHETT & GOLD & E. P. S. Co., Ltd., and H. M. GENESE. Brit. 336,041, Aug. 9, 1929.

Manufacture of dollies for dry cell electric batteries. H. A. BRUMER GRS. Brit. 337,001, April 8, 1929. Mech. features.

Manufacture of dollies for dry cell electric batteries. H. BLACHE. Brit. 337,164, April 2, 1929. Mech. features.

Batteries. OTTO BÖHM, Ger. 513,805, Jan. 17, 1930. Details are given of fastening the electric storage cells together.

Storage batteries. FRITZ HERTSCHEL, Ger. 513,807, Feb. 23, 1930. A continuous process app. for filling dry battery cells with alk. electrolyte, comprising recessed wheels between which curved electrode strips and electrolyte are rolled, is described.

Storage battery. SOCIETÀ ITALIANA BATTERIE ELETTRICHE CREMONA, *Pr* 604,745, Apr 20, 1931. Neg. electrodes are formed of sheets covered with a layer of Zn deposited electrolytically by means of an anode of Zn and in an electrolyte contg alkali silicate.

Storage battery with sulfuric acid electrolyte. T. STAMMAY, *Brit* 336,113, Oct 21, 1929. A H_2SO_4 electrolyte is gelledified by adding 2 parts Na silicate to 5 parts of water, heating the mixt. and then adding it to 2 parts H_2SO_4 dil'd with 5 parts of water.

Storage battery separators. W. R. EDWARDS and R. EDWARDS & CO., LTD. *Brit.* 336,418, Oct 30, 1929. Structural features.

Electrodes for accumulators. PATZ HENNING, *Ger* 513,800, Feb 14, 1930. Details are given of forming electrodes for storage batteries, contg alk. electrolytes, from blocked perforated sheets.

Metallizable oxide electric rectifiers. HERMANN SCHICKNER and WILHELM A. G. *Brit* 330,920, Feb 27, 1929. The oxide layer is coated with graphite, and the latter and a lead to be attached are both smeared with a pasty soft solder compn. contg an acid free deoxidizing agent such as eulophony to which NH_4Cl may be added. Zn is then sprayed onto the oxide layer to form a coating, melt the solder and melt the oxide layer and the lead.

Electric condensers. HERMANN SCHICKNER and WILHELM A. G. *Brit* 330,990, July 18, 1929. Multi part insulator condensers are made so that the outer part forming the insulator is of a material of high mech. and thermal strength such as porcelain, stearite, stoneware or quartz glass, while the material such as glass of the inner condenser part has a high breakdown strength and the highest possible dielectric const. (even if this involves a lessened mech. or thermal strength).

Electrodeposition of copper. A. MOZUR, *Brit* 336,119, Oct 29, 1929. In operations such as plating, manufacturing of electrolytes and refining of black copper, an electrolyte contg. Cupels is employed, e. g., Cu_2O , HCl , KI (2%) and water (1000 parts) may be used with an addn. of 5-10 parts of a colloid such as glue or gelatin or of a substance having capillary activity such as phenol or tannin. A c. d. of 400 amps per sq. meter at 6.5 v. is suitable.

Electroplating materials such as feed-mill screens with metals such as chromium. RICHARD HIRSCHAU (to Swifton Mfg. Co.) *U. S.* 1,702,107, Feb 10. The electrolyte in which an anode is immersed is caused to circulate through the interstices of the material, which is connected as cathode, and at the same time elec. current is passed through the electrolyte to effect plating of the interstices of the material as well as of its exterior surface. App. is described.

Coating objects with chromium. METALS PROTECTION CORPORATION, *Ger.* 513,744, July 3, 1927. See *Pr.* 610,913 (*C. A.* 23, 1905) and *Brit* 277,295 (*C. A.* 22, 2331).

Nickel and chromium coatings on foundation metals such as iron or steel. CUTLER G. FINE and L. CUTLER (to Chemical Treatment Co.) *U. S.* 1,702,082, Feb 10. A heat treated plate of Ni on the foundation metal is provided with an overlying plate of Cr.

Stripping off chromium plating. J. P. MCCARTHY (to Treatment Mfg. Co.) *Brit.* 336,532, March 6, 1929. See *Ger* 511,410 (*C. A.* 23, 1107).

Plating aluminum. LOUIS SCHURER *U. S.* 1,791,612, Feb 10. The surface of an Al article is "passivated" (suitably by treatment with HNO_3), and the prep'd surface is plated with an alloy of Ni and Fe by electrodeposition in a bath contg. sulfates of Ni and Fe and a conc. salt such as Na_2SO_4 , and the coating thus formed is then further plated with a finishing metal such as Ag, Au, brass, Cu, bronze or Ni.

Electrolysis. OTTO LERNER, *Austrian* 129,100, July 15, 1910. In pptg. metals from soln. by electrolysis, the cathodes used are made of Fe (or an Fe alloy) that has been converted superficially into the nitride. A nitrified Fe-Cr alloy is instances. The electrolyte must not be too acid. The ppt'd metals are readily removed.

Electrolytic cells. JOSEPH L. WOODWARD, *Ger.* 513,911, Nov. 20, 1927. An arrangement for insulating the electrodes is described.

Electrolytic cells such as those for producing oxygen and hydrogen. A. R. KNOWLES, *Brit* 335,987, July 6, 1929. Liquid level in an electrolytic gas generator is maintained by flow from a tank, which is fed with electrolyte from gas-washing chambers, and delivery of fresh liquid to the washing chambers is controlled by a float in the tank. Various details of app. are described. *C. I. C. A.* 25, 1107.

Electrolytic production of hydrogen peroxide or other compounds containing active oxygen. FRITZ NOACK, DAVID NITENBERG and GEORGE FETTERBERG (to I. G. Farbenind. A. G.) *U. S.* 1,702,109, Feb. 10. See *Brit.* 316,018 (*C. A.* 24, 1935).

Acetylene. CONSORTIUM FOR ELECTROCHEMICALS INDUSTRIE G. M. & H. Fr. 694,100, Apr. 17, 1930. C_2H_2 is made by submitting a current of CO or CO_2 or both mixed with H or hydrogenated compds. to the action of high tension elec. discharges, preferably under reduced pressure.

Formaldehyde. GUTHRIEFENGSCHMITZ OBERHAUSEN A.-G. Fr. 694,330, April 23, 1930. CH_2O is made by passing a mixt. of CH_4 and CO_2 or gases contg. these through an elec. field produced by an elec. current of high tension and frequency. Cf. C. A. 24, 5241.

Decomposition of water. YUJI MORISUMI. Fr. 694,346, April 23, 1930. An app. is described for decomposing water into H and O, in which water from a chamber can flow into a very narrow space between 2 of 3 coaxial tubes forming a chamber in which the water is vaporized by heat from an elec. heating element disposed in the central tube. The vapor thus produced is introduced into a decompn. chamber in which it is decomposed into its constituents by another heating element.

Electrolytic manufacture of salts. JEAN ALAS. Fr. 694,915, Jan. 15, 1930. Sol. salts of metals placed at the anode are made with the acid salt of alkali metals as electrolyte. The process may be used for making sol. sulfates of Cu, Fe, Al, Ni and Zn and also phosphates, chlorates, perchlorates and persulfates.

Alumina. GINO GALLO. Fr. 694,883, May 2, 1930. Leucite or other silicates contg. Al and K or Na or natural or artificial salts of Al are treated to obtain SiO_2 , Al_2O_3 , and salts of K or Na in a very pure state, by making a soln. of a K- or Na-alum and submitting the soln. to electrolysis to obtain H_2SO_4 at the anode and $NaAlO_2$ or $KAlO_2$ at the cathode. An excess of alkali sulfate is used to facilitate the formation of the corresponding aluminate. The anode may be of Pb and the cathode of Fe.

Zinc. I. G. FARRENTINO, A.-G. Fr. 694,234, April 19, 1930. Max. yields are obtained in the electrolysis of $ZnSO_4$ solns. by maintaining the circulating electrolyte at the most favorable compn. for the working conditions, each individual vat of the app. being charged directly from the circulating vessel.

Electric furnace. EMIL F. RUSS. Ger. 513,762, Aug. 5, 1927.

Electric furnace. ART. GOS. BROWN, ROBERT & CO. Ger. 516,397, June 8, 1928. Means is described for insulating the electrodes against the electrode-moving device.

Electric furnace. SIEMENS-FRANKEWERKE A.-G. FOR KOHLEFABRIKATE. Fr. 694,428, April 25, 1930. Mounting of electrodes is described.

Electric induction furnace. I. G. FARRENTINO, A.-G. (Emil Fendt and Carl Schütz, inventors). Ger. 513,689, Dec. 29, 1928. The spools and rolls, etc., carrying the heating filaments are rendered sufficiently resistant to heat by making them of asbestos-insulated wire impregnated with a cement mass of fluoborate.

Electric induction heating of apparatus such as that for treating oils or further oxidation of lithium. EDWIN F. NORTON (to Ajax Electrothermic Corp.). U. S. 1,791,934, Feb. 10. Coils and elec. connections are arranged so as to provide sections for regulated differential heating.

Electric resistance furnace. W. J. MILLAR and ELECTRIC FURNACE CO., LTD. Brit. 336,097, Oct. 8, 1929.

Electric resistance furnace for carrying out exothermic reactions. ALLGEMEINE ELEKTROLYTISCHES. (Viktor Paschke, inventor). Ger. 513,542, Sept. 27, 1927.

Electric resistance furnace for annealing or other heat treatments. P. CARLBERG. Brit. 336,078, Sept. 23, 1929. Structural features. A cooling medium may be circulated through the resistance, tubular in form, when the heating is completed.

Electric resistance furnace for firing enameled goods. C. L. IRSEN and J. L. McFARLAND (to British Thomson-Houston Co., Ltd.). Brit. 336,168, Dec. 18, 1929. Structural features.

Electric furnace with resistance heaters under the hearth. ALLGEMEINE ELEKTROLYTISCHES. Ger. 516,396, Mar. 1, 1929.

Apparatus for regulating the position of electrodes of tilting electric furnaces. HENNING G. FLODIN. U. S. 1,792,293, Feb. 10. Structural features.

Electric heating elements. C. B. BACKER. Brit. 336,949, June 22, 1929. Mg is applied to a wire in the form of 2 or more longitudinal strips bent in arcuate form so that the strips form a tube about the wire, and the Mg is converted into MgO by the action of steam, a resistance element embedded in the MgO and provided with an outer casing being formed.

Electric smelting oven. EMIL F. RUSS. Ger. 516,190, Mar. 5, 1929. Addn. to 504,419 (C. A. 24, 5241).

Acetylene and hydrogen by electric arc treatment of saturated hydrocarbons.

1. G. FARREYND A-G. Brit. 337,088, Aug. 7, 1929. See Fr. 674,459 (C. A. 24, 2385). Gases such as CH_4 , C_2H_6 , natural gas, distill. gases from coal, shale or peat, cracking gases, oil or tar vapors, or gases contg. nebulized carbonaceous materials such as tar cuts or coal dust (which may be assoc. with H or with small quantities of N, CO or CO_2) are passed through a series of arcs, the lengths of which are adapted to the changing compn. of the gas so that longer arcs are used as the concn. of hydrocarbons decreases. Various details and examples are given.

Thermal decomposition of methane, etc. I. G. FARREYND A-G. Brit. 336,201, July 10, 1929. Gases such as CH_4 , or natural gas are decomposed (for production of C_2H_2 , carbon black and H) by passage through an elec. arc at a temp. of about 2500° or higher, so that the ratio of the quantity of gas (measured in cu. meters per hr.) to the power of the arc, in kilowatts, lies below 0.6.

Iron cores for electrical purposes. WITTEKREKKEKAPPEL G. M. N. H. (Walther Ehlers, inventor). Ger. 513,781, Dec. 21, 1927. Fe powder, ϵ g., obtained by electrolysis, is pressed at 20,000 to 30,000 atm.

Electrical gas purifiers. METALLIGES A-G. Fr. 694,603 and 694,601, April 28, 1930. Electrodes and means for suspending them are described.

The purification of gases and the manufacture of fertilizers. SOCIÉTÉ BRION NITRIL-STRYTHÈSE "SOMERLYN" Belg. 371,622, Aug. 31, 1930. Solid fertilizer is obtained electrolytically. The electrolyte consists of a soln. of some basic salt, preferably liquor contg. CO_2 in some form and org. matter such as ammoniacal liquors from the gasification of coal, liquid manure, etc. A graphite or retort C anode is used. Cf. C. A. 25, 1002.

Electrical purification of blast-furnace gas. SIEMENS-SCHUCKERTWERKE A-G. Fr. 694,210, April 19, 1930.

Alloy armoring for electric cables. W. T. HENLEY'S TELEGRAPH WORKS CO., LTD., and P. ROSSING. Brit. 336,406, Oct. 3, 1929. For reducing armoring losses in single-core a. c. cables, the armoring is formed of an alloy of Al contg. also Mg 0.4-0.5, Sn 0.5-0.6 and Fe about 0.3%.

Loading material for electric cables. STANDARD TELEPHONES & CABLES, LTD., and T. N. RILEY. Brit. 336,620, July 17, 1929. Dissimilar loading materials are used on different portions of a cable, the sections (such as the end sections) to be used at greater energy level being loaded with material of constant permeability over the normal working range (such as Ni 45, Fe 30 and Co 25%) and the intermediate portion being loaded with material of high resistivity (such as Ni 74.5, Mo 3% and the remainder Fe).

Temperature-indicating device for oil-immersed electrical apparatus such as transformers. GENERAL ELECT. CO. and BRITISH THOMSON-HOUSTON CO., LTD. Brit. 336,431, Oct. 14, 1929. Structural features.

Metallic-vapor lamps. N.-V. PHILIPS' GLOBILAMPEN FABRIEKEN. Brit. 336,206, July 5, 1929. Various structural details are described of elec. discharge tubes contg. an alloy or mixt. of metals such as a Cd amalgam, which preferably constitutes an electrode of the tube.

Electric neon lamps. CLAUDE NEON LIGHTS, INC. Brit. 336,956-7, July 18, 1929. Structural and elec. features.

Electric incandescent lamps. N.-V. PHILIPS' GLOBILAMPEN FABRIEKEN. Brit. 335,933, July 1, 1929. Various structural details are given of a lamp having a "concentrated" W filament assoc. with Hg or like material inert to W in a bulb allowing the Hg to circulate and which also may contain gases such as N or Ar. Cf. C. A. 25, 1449.

5—PHOTOGRAPHY

R. P. WIGHTMAN

New processes of color photography. G. GROTH. PAM. Korr. 66, 177-81 (1930).—A review is given of the subject matter of all the patents issued in 1929 for additive (screen), subtractive or direct color processes.

Photography in the infra-red. W. DIETLER. PAM. Korr. 66, 309-15 (1930).—After an historical account of progress in infra red photography by means of filters and suitably sensitizing dyes, from the discovery of chevreau to the present time, D. describes its applications to distance photography, to night-effect exposures (with Kodak or Agfa film), to spectrophotography and photomicrography in the infra-red and to the photography of substances at temps. below a visible red heat. Working prescriptions are given for the infra red sensitizing of ordinary high-speed plates and films and for the

hypersensitizing of the purchaseable Kodak or Agfa infra-red sensitive plates. A suitable light filter, especially suitable for the purpose of the exposure, must be used and development may be carried out by a feeble green light. E. R. BULLOCK

Filter solutions for fluorescence photography. K. WEBER. *Phot. Korr.* 66, 317-9 (1930)—A 2% aq. soln. of NaNO_2 in a thickness of 1 cm. in a glass cell absorbs ultra-violet radiations of 366 m μ almost completely. W. gives comparative illustrations of fluorescence photography with a NaNO_2 filter and with "G. G." (Schott & Gen.) solid glass filters, the latter, though the most suitable of that kind available, do not absorb ultra-violet nearly as satisfactorily as does the former. E. R. BULLOCK

Photographic hardness and absorption measurements of x-rays. L. GREEN AND W. SCHMIDT. *Z. wiss. Phot. (Schweiz. Festschrift)* 29, 129-33 (1930)—By comparing x-ray exposures made through an Al step tablet with those made sensitometrically one may compute from the absorption law the effective wave length of a heterogeneous beam. The method is the same in principle as the ionization method. H. E. S.

The concentration speck or centripetal theory of light action. A. STEGMANN. *Kolloid Z.* 53, 318-22 (1930), cf. *C. A.* 24, 30, 25, 44—The theory is applied to the mechanism of the action of light on photographic plates. ARTHUR FLEISCHER

Sensitiveness of photographic dry plates. M. MIYATA. *Mem. Coll. Eng., Kyoto Imp. Univ.* 6, No. 2, 113-72 (1930)—A possible additional cause of the sensitivity of the dry plate is presented. It is found that AgCl is always present with AgBr at its pptn. and it is presumed that the AgCl grains are the sensitive centers in the AgBr crystals. Further expts. show that AgBr and AgI in photographic emulsions act as both chem. acceptors and optical sensitizers for AgCl nuclei. Ripening of an emulsion is the result of the adhesion of AgCl and AgBr crystals sep'd. at pptn. by gelatin. L. A. STAIN, JR.

The accumulation effect of photographic plates and its influence on the accuracy of photographic photometry. N. BARABASCHOFF AND B. SEWEYKIN. *Z. wiss. Phot.* 28, 333-40 (1930)—The accumulation of photographic density (growth of the latent image) is real and must be taken into account in photometry. It is advisable to develop plates at least 2 days after exposure, otherwise systematic errors which influence the results are introduced. Old plates show this accumulation effect to less extent than new plates. A. P. H. TRIVELLI

Reproduction of negatives. P. HANNYER. *Camera (Luzern)* 9, 155-7 (1930)—A AgBr plate or film is bathed in 4% $\text{K}_2\text{Cr}_2\text{O}_7$ soln., dried and exposed under the negative to be copied until a faint brown image is visible. The duplicate is then washed and finally developed by ordinary light in a regular photographic developer. The unexposed portions develop, giving a duplicate negative. The contrast can be controlled by the length of exposure or by a subsequent flash exposure. M. W. SEYMOUR

Dyes unstable to light are sought. H. HERNSPORR. *Camera (Luzern)* 9, 160-4 (1930)—The theory and history of color printing processes by means of bleach-out dyes are briefly described. The problems to be solved in the improvement of the bleach-out process are said to be increasing the sensitivity and working out a good means of fixation. A process useful practically requires plates or papers that will keep for at least 6 months to a year. M. W. SEYMOUR

Role of dyes in the progress of photography. A. SEWEYKIN. *Photo-Revue* 42, 353 (1930)—In the application of dyes to photography the following uses have been found: (1) chromatic sensitizing of photographic negative materials, (2) photographic reproduction of colors, (3) lighting of the darkrooms and the chromatic desensitizing of plates, (4) prep'n. of antihalo plates, (5) tinting of plates and films for transparencies, (6) prep'n. of colored motion pictures and (7) production of photographic images with color compounds. Most of the dyes which are used for the chromatic sensitizing of photographic materials belong to the following types: (1) d-phenylmethane, (2) triphenylmethane, (3) azine, (4) phenylacridine and (5) quinoline. A description including the spectral region which they sensitize is given for several dyes. *Ibid.* 375 (1930)—The cyanine dyes used as sensitizers for photographic materials are classified into 4 groups: (1) the cyanines, (2) the isocyanines, (3) the pinacyanols and (4) the diacyanines. A brief description is given for several dyes belonging to each class. The use of desensitizing dyes after exposure in order to permit inspection with a bright light during development was introduced by Lüpke-Cramer with the discovery of phenosafranine and has led to a study of the desensitizing properties of other dyes. Certain dyes which sensitize Ag chloride and bromide emulsions were found to desensitize Ag iodide emulsions. *Ibid.* 43, 5-7, 23-4 (1931)—No direct relation exists between the constitution of a dye and its desensitizing properties. All of the safranines, however, which contain amino groups have desensitizing properties. The expts. of Lüpke and Seyewitz have indicated that the desensitizing action of phenosafranine is

due to the formation of an adsorption complex between the AgBr and the dye which is not sensitive to light and is slowly destroyed by washing. LÜPPO-CRAMER considers that the desensitizing action is due to the oxidation by the dye. Dyes are used in the prepn. of films and plates for the suppression of photographic halation due to reflection. In this case the dye is incorporated in the subcoating between the sensitive emulsion and its support or is applied as a dyed gelatin backing. One of the most important uses of dyes has been in the photographic reproduction of color by the indirect method. A practical discussion with formulas is given for this process on the selection of the dye and the dyeing of the microchromes. H. D. RUSSELL

Acceleration of development by neutral salts and dyes. LÜPPO-CRAMER *Phot. Korr.* 66, 253-6(1930).—L. C. for the first time gives density curves to illustrate the phenomenon of acceleration of development by the addn. to the hydroquinone or other developer of large amts. of neutral salts or relatively very small quantities of certain basic dyes. Plates of very contrasty emulsion (not further described) were used, and the standard developer was hydroquinone with sulfite, K_2CO_3 and KBr. With these the effect of KNO_3 (10%) or of pinakryptol green (0.005%) is very conspicuous. As regards theory, L. C. remarks that the action of the neutral salts and also, though there are some exceptions and contradictions in this case, that of basic dyes, reveals their coagulating action on such suspensions as colloidal Ag. The exceptional difference between the action of pinakryptol green and pinachrome, the latter producing log and a relatively flat gradation, is illustrated. F. R. BULLOCK

Fine grain developers and their application to spectrography. DALLIN VIERST *Naturwissenschaften* [N. S.], 6, 38-9(1929). *Proc. 7th Intern. Congr. Phys.* 1928, 255-7.

The role of sulfite in photographic developers. J. RAYMOND *Chimica* (Luzern) 9, 125-32, 164-7(1930).—Evidence is adduced from the literature and from original exptl. work to show that the protective action of sulfite against the aerial oxidation of hydroquinone developers can be explained by the following series of reactions: (1) $C_6H_4(OH)_2 + O \rightarrow O-C_6H_4-O + H_2O$, (2) $3Na_2SO_3 + H_2O \rightarrow 2Na_2SO_4 + S + 2NaOH$, (3) $S + Na_2SO_3 \rightleftharpoons Na_2S_2O_5$, (4) $O-C_6H_4-O + Na_2SO_3 + H_2O \rightarrow (HO)_2C_6H_4SO_3Na + NaOH$, (5) $(HO)_2C_6H_4SO_3Na + NaOH \rightarrow Na_2SO_3 + (HO)_2C_6H_4SO_3H$, (6) $(HO)_2C_6H_4SO_3H + O + NaOH \rightarrow (HO)_2C_6H_4SO_3H + H_2O$. The hydroquinone-sulfite complex previously assumed by Mees and Dyer (*Bull. J. Phot.* 1912, 441-3) to explain the fogging action and rational cond. of hydroquinone developers with a deficiency of sulfite is believed to be the hydroquinone monothiosulfonic acid formed in equation 4. The occurrence of hydroquinone disulfonic acid in oxidized hydroquinone developers is explained by assuming that the monothiosulfonic acid, formed in equation 6, undergoes the same series of reactions as the hydroquinone itself. M. W. SYMONS

A critical examination of the Hübl system of color sensitometry. E. HEINENBERG AND M. BILTZ. *Phot. Korr.* 66, 169-76, 198-200(1930).—H. and B. find that the color reproduction of materials calcd. according to the Hübl system agrees with the true color reproduction within an error of at most 20%. The variation of emulsion gradation (gamma) with the filter (wave-length range) needs to be taken into consideration; the underexposure of cut of the filters, and the variation in color sensitivity of the emulsion within the limits of transmission of any one of the filters are, on the other hand, generally practically negligible. E. R. BULLOCK

Solarization. IV. Dependence of solarization upon the preparation of emulsions. H. ARENS. *Z. Wiss. Phot.* 28, 341-7(1930); cf. C. A. 25, 601.—An emulsion freshly prep'd from a peptized Ag compd. solarizes a little or not at all. After some time the power to solarize increases. This increase is greater the less the concn. of Ag ions. If they are removed as far as possible by KBr, in a few seconds the emulsion acquires the power to solarize very strongly. This process is irreversible. A. P. H. DEVELL

Constancy of the amount of light in magnesium foil sensitometry. G. SCHWARTZ AND F. URRACH. *Phot. Korr.* 66, 261-4(1930).—Using a micro-burner for igniting the Mg and a Rb and a K photoelectric cell for measuring the light, S. and U. conclude from the results of their expts. that Mg foil as prescribed by Rieder (cf. C. A. 23, 4153) gives a practically const. source of light sufficient for this kind of sensitometry. It is necessary, however, by const. manipulation to diminish fluctuations in the light and also to take the mean of several measurements to bring the residual accidental error below that of the Eder-Hoch sensitometer. E. R. BULLOCK

Physics of the washing-out of hypo from prints. R. E. LIESegang. *Phot. Korr.* 66, 324-6(1930).—Expts. with ferrous NH_4 sulfate and with $CuCl_2$ suggest that, in consequence of the high osmotic pressure, hypo is removed in the washing process

from fixed prints more quickly than would be expected on the basis of simple diffusion.

E. R. BULLOCK

Displacement of silver in partly swollen gelatin layers. E. WAGNER AND K. SCHAUUM. *Z. wiss. Phot.* 28, 325-7(1930)—A drop of water on a photographic plate or film causes a spot after development, in which the edges have higher densities than the center. This is due to a displacement of the gelatin contg. Ag halide grains to the edge of the wetted spot. The extent to which the real Ag density or addnl. light reflections cause the dark edge can be investigated by detg. the edge density after immersing the surface of the film or plate in glycerol. It is shown that even in that case the edges have a higher density, because of migration of Ag halide to the edge of the spot.

A. P. H. TRIVELLI

Hydrolyses occurring in washing of the gelatin-silver nitrate system. H. H. SCHMIDT AND P. PRETSCHNER. *Z. wiss. Phot.* 28, 328-32(1930)—Lippmann-Cramer had observed that in the gelatin system plus AgNO_3 , hydrolysis takes place. The authors made a quantitative investigation of this phenomenon as a support for their theory that during pptn. of Ag halides in gelatin, some Ag is formed by the action of AgNO_3 on gelatin. The hydrolysis of AgNO_3 in the presence of gelatin gives Ag_2O and HNO_3 , which Ag_2O is retained by gelatin.

A. P. H. TRIVELLI

Stability of the latent image of a silver iodide emulsion toward dichromate-sulfuric acid. H. H. SCHMIDT. *Phot. Korr.* 66, 315-7(1930)—S distinguishes 2 latent light images, one occurs in AgI-gelatin emulsion formed with excess of sol. iodide, the other in emulsions formed with excess of Ag salt. The former latent image is found to be very unstable and readily destroyed by solns. of oxidizing agents, while the latter is stable even toward dichromate-sulfuric acid. The latter is, moreover, not susceptible or intensification by nucleus denudation, but is on the contrary, apparently somewhat diminished for phys. development, by the action of an intermediate bath of KI.

E. R. BULLOCK

Photographic effects of γ -rays (ROGERS) 3. Manufacture of photographic lenses (GULFORD) 19. Photographic sound records (Brit. pat. 330,550) 18.

Color photographs. LOUIS DUFAY (to Old Jewry Trust, Ltd.). U. S. 1,792,418, Feb. 10. In a process for the reproduction of color photographs on a transparent film bearing only one sensitive layer on each of its faces, the sensitive layers are impregnated with tartrazine, corresponding neg. plates are simultaneously applied on each of the sensitive layers, and the film and neg. plates are exposed to light, and the film is subsequently washed to eliminate the tartrazine and the layers are differently colored.

Color photography. J. N. GOLDSMITH, T. T. BAKER and SPICEAS, LTD. Brit. 337,073, July 30, 1929. A film of material such as celluloid, viscose or the like, to which a multicolor screen is to be applied by dyeing, applying a resist pattern, bleaching and redyeing the parts unprotected by the resist, is preliminarily coated with an isolating layer of substantially waterproof character, such as a synthetic resin or cellulose acetate varnish, and a surface layer of material such as collodion or cellulose prepolymer from viscose which is receptive to dyes.

Color photography. T. T. BAKER, A. B. KLEIN and COLOUR SNAPSHOTS (1928), LTD. Brit. 337,057, July 27, 1929. A "tripack" is formed with a green-sensitive front layer, a blue-violet sensitive intermediate layer, a red sensitive rear layer and (in order to protect the rear layer from green and blue violet light) filters preferably comprising a green absorbing filter between the front and intermediate layers and a blue-violet-absorbing filter between the intermediate and rear layers. Various details of manufacture are described.

Color photography. T. T. BAKER. Brit. 337,040, July 25, 1929. In a process of obtaining prints in natural colors in which a light-sensitive layer coated over a multicolor screen on a transparent support is printed through the support from a negative associated with a multicolor screen (the image and screen being subsequently transferred to an opaque support with the screen uppermost), the screen is printed on a celluloid, cellulose acetate, regenerated cellulose, collodion or gelatin layer attached by an adhesive such as gum dammar to a translucent stripping support such as celluloid or parchmentized paper. The final support may be paper, a cellulose acetate film contg. baryta or material such as is described in Brit. 282,980 (C. A. 22, 3848) which is coated with an adhesive such as gum arabic. Brit. 337,041 describes a process in which, to avoid the moiré effect in copying by projection from a master cinematograph film, a sheet of plane glass scored or indented with lines is placed in front of, or behind, the

projection lens. The lines may be formed as concentric circles about 0.125 in. apart. Cf *C. A.* 24, 5245.

Color photography. I. G. FARBENIND A.-G. *Fr.* 604,031, April 16, 1930. Color images are obtained on films with microscopic lenticular gratings by producing, during the taking of the view, all the images of the colored screen in the layer of emulsion so that they coincide with the central axis of the corresponding lenticular elements.

Photographic layers. PHILIPPE FRANGIALLI and GEORGES H. A. FREYSS. *Fr.* 605,006, Aug. 2, 1929. Layers sensitive to light are made on suitable supports by using as sensitive material diazonium salts of aromatic amines which contain the imidazole group, simple or substituted, mixed with an azo compound of phenolic or aminophenolic properties and adding, according to the case, an acid which does not deteriorate the support. Thus, Schaeffer acid and oxalic acid are added to a soln. of diazotized *N*-methyl *N*-methylaminobenzimidazole and the mixt. is used for coating supports which are then dried. Other examples are given.

Photographic surfaces. JOS-PE FARBENPHOTO G. M. B. H. (Gustav Koppmann, inventor) *Ger.* 513,277, Feb. 17, 1929. See *Brit.* 327,283 (*C. A.* 24, 4996).

Photographic films from mixed cellulose esters. C. S. WEBBER and C. J. STAUD (to Kodak, Ltd.) *Brit.* 336,151. Mixed esters are employed which contain an acyl radical and a radical of an acid containing a hydroxy group in the α -position or a ketonic group in the α - or γ position and which are sol. in acetone or its homologs or in 50–75% aq. alc. soln. and these esters are dissolved in a mixt. of ethylene chloride and a monohydroxy aliphatic alc. containing 5 C atoms or less.

Combined cinematograph and sound record films. T. T. BAKER. *Brit.* 333,899, July 1, 1929. Before application of the rescan to the surface of the film, the narrow marginal strip which is to serve as the sound track is protected from the dyes in the rescan by application of a layer of protecting varnish such as a soln. of bitumen in *Ccll*. The rescan is then applied to the whole surface of the film, followed by treatment with a varnish-removing solvent, so that a strip of clear uncolored film is provided to receive the sensitive emulsion upon which the sound record can be photographed in black and white.

Fire-resistant motion picture films. ARTHUR ARENT (to Arthur Arent Laboratories, Inc.) U. S. 1,792,457, Feb. 10. Films are treated with a soln. prepd from urea and SbCl_3 or from other suitable acid-binding org. compd. which does not promote inflammability and a suitable acid-yielding fire retardant metal salt, applied in a substantially non aq. solvent such as EtOAc .

Photographic developers. I. G. FARBENIND A.-G. (Gustav Reddelien and Werner Müller, inventors) *Ger.* 516,000, Jan. 8, 1930. Addn. to 454,830. *N*-Monohydroxypropyl and *N*-monohydroxybutyl derivs. of *o*- and *p*-aminophenol and their nuclear substitution products not containing NO_2 groups are used as the active constituents of developing solns. for Ag halide layers. A suitable soln. comprises *p*-(hydroxypropylamino)phenol- HCl 1, Na_2SO_3 9 and K_2CO_3 8 g. in H_2O 200 cc. Cf *Ger.* 467,818 (*C. A.* 23, 572).

Transforming photographic silver images. F. LIEBIG. *Brit.* 335,930, July 2, 1929. The opaque substances such as AgBr or Ag are converted into transparent complex compds. having strong mordanting properties by treatment with a reagent such as an iodide, urea, thiocyanate or thiocarbamide and other substances such as bisulfites may also be added which react with the complex forming reagent. Various details and examples are given.

6—INORGANIC CHEMISTRY

A. R. MIDDLETON

The isomerism of radicals. CLEMENT DUVAL and MME. CL. *Compt. rend.* 191, 843–5 (1930), cf *C. A.* 24, 563—A special study with isomer ions $[\text{CoSO}_4(\text{NH}_3)_5]^-$ and $[\text{CoSO}_4(\text{NH}_3)_4]^{2-}$. Persulfatepentammine sulfate has a d. of 2.1 and a mol. susceptibility of -588×10^{-6} . Sulfatepentammine sulfate is red, has a d. of 1.7 and a mol. susceptibility of -212×10^{-6} , measured at 15°. Both give the same product on hydrolysis in dil. soln. GERALD M. PETTY

The coordination value of multivalent negative radicals. PAUL PFIEFFER. *Z. anorg. allgem. Chem.* 192, 366–8 (1930)—If it is assumed that Co has a const. coordination no. of 6, certain bivalent radicals sometimes have a coordination value of 1 and sometimes of 2, as follows: $[(\text{H}_2\text{N})_2\text{Co}(\text{CO})_2]\text{X}$ and $[(\text{H}_2\text{N})_2\text{Co}(\text{CO})_2]\text{X}$. These can be satisfactorily explained by the following distribution of charges on disocn.: $^{++}[(\text{H}_2\text{N})_2\text{Co}(\text{CO})_2]^{--}$.

$N)_3Co-O-CO-O^-] + X^- \cdot [(H_2N)_3Co \begin{array}{c} \diagup O \diagdown \\ \diagdown O \diagup \end{array} CO]^+ + X^-$. The following general

statement can be made bivalent acid radicals take 2 coordination positions in the uncharged condition and only 1 coordination position when carrying a single neg charge If 2 neg charges are carried, then the coordination no of the radical is zero

II. STORATZ

The mechanism of precipitation changes. I. The reaction between lead chloride and ammonium chromate. Z KARAOGLANOV AND B SAGORTSCHEV. *Z. anorg allgem. Chem.* 194, 151-8 (1931) — $PbCrO_4$ pptd in the presence of HCl contains Cl , whose quantity varies with the concn of Cl^- , the time required for complete pptn and the temp. The Cl held by the ppt does not appear to be volatile at high temps. When HNO_3 is present no nitrate could be detected in the ppt. No dichromate was found in either case. The Cl is believed to be held as $Pb_2Cl_2CrO_4$ (cf. *C. A.* 24, 4690). The results are reproducible only with carefully controlled conditions. II. The reaction between sulfuric acid and lead bromide. *Ibid.* 195, 105-12 — Pptn of $PbCrO_4$ with $PbBr_2$ gives the same results as with $PbCl_2$. The reproducibility of the results depends on the length of time the ppt stands in the soln. The longer the time the less the Br in the ppt. III. Further investigations of the reaction between sulfuric acid and lead bromide. *Ibid.* 113-20 — The decrease in Br content of the ppt, which occurs when the ppt is not immediately filtered has been studied. There appears to be a meta-stable Br -contg product which is not a mixed crystal, solid soln or abs compd but rather a complex of the type $(PbBr)_2SO_4$ which decomposes on standing in soln to give $PbSO_4$ and $PbBr_2$. J. B. AUSTIN

Investigations on cerium. LEON LORIE. *Ann chim* [10], 14, 407-60 (1930) — Ceric salts (valence 4) dissolve in alkali carbonates to form complex salts. The ceric carbonate of Na , $[Ce(CO_3)_2 (H_2O)_2]Na$, $10H_2O$, has been sep'd, as yellow crystals and analyzed. It decomposes in H_2O to cerhydroxy carbonate, $CeCO_3(OH)$. Al though K , NH_4 , Li and Tl form similar compds, they could not be crystd. The Na salt is isomorphous with the Na thorcarbonate, $[Th(CO_3)_2 (H_2O)_2]Na$, $10H_2O$. Tl , Zr and Ti salts show similar reactions. Cerous tartrate (valence 3), $Ce_2(C_4H_4O_6)_3 \cdot 6H_2O$ or $Ce_2(C_4H_4O_6)_3 \cdot 4\frac{1}{2}H_2O$, dissolves in halogen acids to form complex acids as cerous chlorotartarate, $Ce_2(C_4H_4O_6)_3 \cdot 3HCl$, $14H_2O$, and cerous bromotartarate, $Ce_2(C_4H_4O_6)_3 \cdot 3HBr$, $14H_2O$. With tartaric acid it forms cerotartaric acid, $Ce(C_4H_4O_6)_2 \cdot 11H_2O$ which is able to form the Na cerotartarate, $Ce(C_4H_4O_6)_2Na$, $8H_2O$. In alk tartrates it forms ceric compds, which are subsequently reduced by the tartrate to cerous compds. With excess O unstable perceric compds (valence 6) form. Cerous tartrate dissolves in alkalis to form salts as K ceritartarate, $2CeO(C_4H_4O_6)_2K_2$, $C_4H_4O_6K_2$, $11H_2O$. These are red crystals sol in water, bases and acids. With Na salicylate cerous salts give perceric salicylate, $Ce_2O_3(C_6H_4O_2CO_2O)_2$, $6H_2O$. A. LE V.

Germanium. VIII. The action of ammonia on germanium tetrachloride: germanium imide. JOHN S. THOMAS AND WM. PUGH. *J. Chem. Soc.* 1931, 60-71; cf. *C. A.* 24, 2391 — The principal object of this investigation is to det. whether $GeCl_4 \cdot 6NH_3$ is a pure compd or whether it is a mixt. of $Ge(NH_3)_6$ and NH_4Cl . Since $Ge(NH_3)_6$ is readily hydrolyzed when exposed to traces of H_2O vapor, an all glass app. is described in which $GeCl_4 \cdot 6NH_3$ can be prep'd, washed with liquid NH_3 , dried and weighed without opening at any stage. Thus, $Ge(NH_3)_6$ of 97.2% purity was prep'd. By means of liquid NH_3 , $GeCl_4 \cdot 6NH_3$ was broken up into 1 mol $Ge(NH_3)_6$ and 4 mols NH_4Cl , thus it is a mixt. rather than a mol compd of $GeCl_4$ and NH_3 . The existence of $Ge(NH_3)_6$ as an intermediate product of the action of NH_3 on $GeCl_4$ above -20° could not be confirmed. When HCl is passed over $Ge(NH_3)_6$, the addn compd $Ge(NH_3)_6 \cdot 2HCl$ is first formed, this then combines with 4 mols of HCl to form $GeCl_4$ and NH_4Cl . The action of heat (up to 300°) on mixts. of $GeCl_4$ and NH_4Cl was studied. The re-

sults are explained on the assumption that $HN \begin{array}{c} \diagup N \diagdown \\ \diagdown N \diagup \end{array} Cl$ is formed



J. BALOZIAN

The chemistry of germanium. ROBERT SCHWARTZ, P. W. SCHENK AND H. GIESF. *Ber.* 64B, 362-8 (1931), cf. *C. A.* 24, 3189 — From vapor pressure and x ray measurements hydrated GeO_2 is found not to exist. GeO_2 m. $1115 \pm 3^\circ$. The prep'n of GeO_2 from $Ge(OEt)_4$, of Ge_2OCl_4 and of $Ge(SO_4)_2 \cdot GeO_2$ is described. V. F. II

Nitrogen compounds of germanium. I. The preparation and properties of ger-

manic nitride. WARREN C. JOHNSON *J. Am. Chem. Soc.* 52, 5160-5 (1930).—Ge was heated in NH_3 at 650° to 700° , yielding Ge_2N_4 , a light brown powder. The Ge_2N_4 is reduced by H_2 at 700° , reacts with O_2 at 850° and with Cl_2 at 600° to 700° ; decomposes at 900° to 1000° and resembles the corresponding compd. of Si in its high stability.

GERALD M. PETTY

Studies on heteropoly acids of germanium. I. Germanomolybdic acid. CHARLES G. GROSSCUP *J. Am. Chem. Soc.* 52, 5154-44 (1930).—A heteropoly acid of Ge and Mo having the formula $\text{H}_{12}[\text{Ge}(\text{MoO}_4)_4]\text{aq}$ was prepd. The coeff. for "aq" has an apparent max. of 26. The use of MoO_3 for the colorimetric detn. of Ge is suggested.

GERALD M. PETTY

Chlorogermanates of alkaloids and chlorogermanates of cesium. ARAKEL TCHAKIRIAN *Compt. rend.* 192, 233-4 (1931). GeCl_4 forms with alkaloids, as quinine and pilocarpine salts of 2 types. GeCl_4 , A 2HCl (dibasic alkaloid) and GeCl_4 , A HCl (monobasic alkaloid) in which A represents the mol. of the alkaloid. The method for the prepn. of the salts and their properties are described. If to a soln. of GeCl_4 strongly acidified with HCl a soln. of CsCl is added, a sparsely sol. white microcryst. ppt. is obtained. Its formula is $\text{GeCl}_4 \cdot \text{CsCl}$. An analog is obtained with RbCl , but not with LiCl , NaCl or KCl . At 100° in an atm. of dry HCl the Cs salt liberates the GeCl_4 . The stability of the new compd. is such that at the temp. of fusion in an atm. of dry HCl GeCl_4 is not liberated but GeHCl_3 is formed instead. It is possible, therefore, that the Ge exerts its tri-coordinated valence and that the formula of the compd. is $[\text{GeCl}_4]\text{Cs}$.

B. S. LIVING

Sulfides of the rare earths. W. KLEMM, K. MEISEL AND H. U. VOGEL. *Z. anorg. allgem. Chem.* 190, 123-44 (1930).—The disulfides of La and Ce, and the sesquidisulfides of Sc, Yt, La, Ce, Pr, Nd, Sm, Gd, Dy, Yb were prepd. by treating the anhyd. sulfates or chlorides in a stream of dry H_2S at elevated temps. The d.s., colors, magnetic susceptibilities and x-ray cryst. structures were studied. On the basis of the color and x-ray detns. the sesquidisulfides may be divided into 2 series, the first of which (La Dy) shows a regular variation in the mol. vol., while the second (Dy, Yb, Yt, Sc) has an irregular variation. The magnetic properties indicate that the sesquidisulfides are salts. The disulfides were proved to be of a polysulfide nature.

L. L. QUILT

The sulfates of zirconium and hafnium. GEORG V. HEVESY AND TRIKA CREMER. *Z. anorg. allgem. Chem.* 195, 339-44 (1931).— $\text{Zr}(\text{SO}_4)_2$ and $\text{Hf}(\text{SO}_4)_2$ prepd. from the tetrahalides or the oxides, always contain an excess of H_2SO_4 . This is the reason why the at. wts. of Zr and Hf, as detd. by the sulfate method, are too small. The last traces of H_2SO_4 are removed only at temps. at which the sulfates decomp. noticeably. The relation between the decompn. pressure of the sulfates and the temp. was detd. between 550° and 650° . $\text{Zr}(\text{SO}_4)_2$ has the higher decompn. pressure.

L. K.

Properties of cuprous oxide. O. A. ARWAS. *Naturwissenschaften* 19, 133-4 (1931).—The sp. resistance of artificial Cu_2O depends on various factors, such as the pressure of the surrounding gas: a decrease in air or H_2 pressure gives a decreasing resistance, a decrease in O_2 pressure causes an increase in resistance, while N_2 or Ne has no effect. The resistance is increased several hundred % by increased gas removal from the substance in high vacuum. Cu_2O has a reversible transition point at 56° , above which the previously transparent substance becomes opaque. This transition point is less distinct after removal of gas, O_2 addn. causes a reappearance of it. The thermal expansion coeff. of Cu_2O depends likewise on the condition of the Cu_2O lattice. Fresh Cu_2O has practically no expansion up to 56° ($\alpha = 0.00000005$), while above 56° it contracts. On cooling it contracts more until 56° is reached, the length changing irreversibly. If previously the gas has been extd. in *vacuo* the coeff. is 0.000001, changing little at 56° . It is concluded that the phys. properties of Cu_2O depend largely on substances dissolved in (O_2) or adsorbed on (H_2) the lattice. For the change in transparency, formation of $\text{CuO} + \text{Cu}$ is believed to be responsible, although this reversibility remains unexplained. The photo effect of Cu_2O can also be related to the gases moving about in the lattice by the theory of phosphoric centers. Other oxides as those of Ni, Sn and Zn (Skaupv., *C. A.* 16, 2111), seem to have similar qualities.

B. J. C. VANDER HORVEN

Structure of silicates. W. L. BRAGG. *J. Soc. Glass Tech.* 14, 295-305 (1930).—See *C. A.* 24, 5553.

H. F. KRIEGER

Structural principles of compounds of boron and hydrogen. I. WIBERG. *Z. anorg. allgem. Chem.* 187, 302-8 (1930). cf. *C. A.* 23, 2897.—Polymers against Hellriegel (*C. A.* 24, 2549). The conception of a "special B linking" to explain the existence of the different B hydrides is unnecessary. From its position in the periodic table it is reasonable to regard B as quinquivalent toward H. As proof that Grimm's hydride

law may be applied to B the similarity between the properties of diborane, B_2H_6 , and C_2H_4 , is pointed out. This indicates that B is quinquivalent in this compd., the 2 B atoms being joined by a double linking. B. C. A.

Preparation of anhydrous stannous chloride. HENRY STEPHEN. *J. Chem. Soc.* 1930, 2766-7.— $SnCl_4$ (long needles) may be more simply prepd than as previously described (*C. A.* 19, 3261) by treating $SnCl_4 \cdot 2H_2O$ (1 mol) with Ac_2O (2 mols.), washing the anhyd. salt, sepg. with dry Et_2O until free from $AcOH$, and crystg. from $AcOH$ contg. a little Ac_2O . J. BALOZIAN

Preparation and purification of the triiodides of antimony and arsenic for use in immersion media of high refractive index. H. G. FRISK. *Am. Mineral.* 15, 263-66 (1930). SbI_3 and AsI_3 were synthesized by gently heating a mixt. of the metals with I in a large test tube. The crude iodides were purified by recrystn. from CS_2 . In this step a small excess of the metals was necessary for combination with free I in the soln. A second recrystn. was made further to purify the compds. In prep. melts from the iodides and piperine it was found essential to avoid heating the mixts. above 135° to prevent decompn. A. M. BRANT

Nitrogen triiodide. HERBERT W. CREMER AND DONALD R. DUNCAN. *J. Chem. Soc.* 1930, 2750-4.—It has been found that NI_3 is formed by the action of NH_3 on the less stable alkali diiodomides (i. e., those with high dissociation pressures). With the K salt as typical, the reaction is $3KIBr + 4NH_3 = 3KBr + 3NH_4Br + NI_3$; the mechanism of which most probably is: $3KIBr = 3KBr + 3IBr$, $3IBr + NH_3 = 3HIBr + NI_3$, $3HIBr + 3NH_3 = 3NH_4Br$. As found by Dobbin and Masson (*Ibid.* 49, 848 (1886), *J. prakt. Chem.* 31, 37 (1885)), the more stable diiodomides such as the tetramethylammonium and trimethylsulfonium salts (having very low dissociation pressures) do not react with NH_3 to form NI_3 , but form additive compds. with 2 mols. of NH_3 . J. BALOZIAN

A comparison of the action of halogen salts of iron on phenyl magnesium bromide and phenyl zinc chloride. G. CHAMPESTER. *Bull. soc. chim.* [4], 47, 1131-7 (1930).—Anhyd. halogen salts of Fe did not give stable organometallic compds. on reacting with $PhMgBr$. The 3 reactions obtained were: (1) $6C_6H_5MgBr + 2FeCl_3 \rightarrow 3(C_6H_5)_2Fe + 2Fe + 3MgBr + 3MgCl_2$, (2) $2C_6H_5MgBr + 2FeCl_3 \rightarrow (C_6H_5)_2Fe + 2Fe + 2MgBr + 2MgCl_2$, (3) $2C_6H_5MgBr + FeCl_3 \rightarrow (C_6H_5)_2Fe + Fe + MgBr + MgCl_2$. $PhFeI$, a new organometallic compd., was prepd. according to the reaction $2C_6H_5ZnCl + 2FeI_3 \rightarrow 2C_6H_5FeI + ZnCl_2 + ZnI_2$. On hydrolysis, the reaction was $2C_6H_5FeI + 2H_2O \rightarrow 2C_6H_5Fe(OH)_2 + FeI_3$. GERALD M. PERRY

The action of halogens on ethyldiiodostibine. Antimony bromodiiodide. ROBERT E. D. CLARK. *J. Chem. Soc.* 1930, 2737-9.—The action of I, Br and Cl on ethyldiiodostibine (I) is studied. SbI_2 (1.2 g., red needles, m. $168-170.5^\circ$) may be prepd. by adding a soln. of I (1.2 g.) in $CHCl_3$ (10 cc.) by drops to I (0.75 g.) in a min. of $CHCl_3$, refluxing for a few min. and cooling; the SbI_2 sepg. out. Antimony bromodiiodide, $SbBrI_2$ (long yellow needles (2.9 g.), m. 88°) is formed when Br (1.6 cc. dissolved in $CHCl_3$ (5 cc.)) is added in drops to a soln. of I (1.2 g.) in a min. of $CHCl_3$ (about 15 cc.), cooled in a freezing mixt., refluxed for a few min., cooled to 0° , the $CHCl_3$ decanted off, the mixt. treated with xylene until free from I, and the residue recrystd. from $CHCl_3$ or light petroleum (b. p. $80-100^\circ$). As typical of the reaction between Cl and I, into a soln. of I (10.0 g.) in warm $CHCl_3$ (10-15 cc.), Cl_2 (530 cc. in all, calcd. 600 cc.) is passed at the max. rate (in this case 35 cc. per min.), at which absorption is complete for the temp. employed ($25-30^\circ$), the mixt. is boiled filtered and cooled, red SbI_2 and its yellow metastable modification alone being produced. J. BALOZIAN

Chloramine, its preparations, properties and uses. W. HIRSCHMANN. *Western Construction News* 5, 531-2 (1930); *U. S. Pub. Health Eng. Abstracts* 11, W. 21 (Feb. 21, 1931).—The chemistry, properties and application to water sterilization of the chloramines (NH_2Cl , $NHCl_2$, and NCl_3) are discussed. C. R. FELLERS

A note on the preparation of hydriodic acid. T. H. WHITEHEAD. *J. Chem. Education* 8, 541-2 (1931).—In the prepn. of HI from H_2S and I, it has been found that the formation of I_2 on distg. the HI soln. can be prevented if the mixt. is dialyzed through a collodion membrane for 48 hrs. This removes the colloidal S which seems to cause oxidation of the HI. R. H. CARTER

The action of bases on non-metals. A. LAURENCE CURL. *J. Chem. Education* 8, 490-7 (1931).—Most of the non-metallic elements with the exception of N and the rare gases react with the alkali hydroxides. In general the reactions take place in accordance with the type $Cl_2 + 2NaOH \rightarrow NaCl + NaClO + H_2O$. The exact mechanism of this type of reaction is not known or discussed. Similar reactions are obtained with the alkali amides which are analogous to the hydroxides in the NH_3 .

system of compds. In NiH_2 the polysulfides, polyphosphides, etc., are more stable than in water, but the ammonio salts are less stable than the corresponding aquo salts.

R. H. CARTER

Course of the reactions between graphite and oxides as well as between heavy metal carbides and oxides. OSKAR MEYER *Arch. Eisenhüttenw.* 4, 103-8(1930) — These reactions were studied by detg. the course of vapor pressure-temp. curves. In the case of SiO_2 and graphite, the partial pressure of CO (p_{CO}) reaches 20 mm. at 1400° . MnO is reduced by graphite at lower temps. p_{CO} having attained about 130 mm. at 1200° . Curves and similar data are given for argillaceous earth and graphite and reactions between the carbides of Si , Mo , W and Cr and the oxides of Mn , Fe , Cr and Si .

H. STORITZ

The interaction between iodine pentoxide and nitric oxide. MADHAVLAL SUKHLAL SHAH AND TRAMRAKAL MOHANLAL DSA *J. Chem. Soc.* 1931, 32-6 — The course of the reaction at 120° between I_2O_5 and NO was followed by measuring (a) the vol. of NO used, (b) the amt. of I_2 produced and (c) the amt. of nitrate formed. The reaction may take place according to either (I) $10\text{NO} + 2\text{I}_2\text{O}_5 = 2\text{I}_2 + 5\text{N}_2\text{O}_5$, and subsequently $5\text{N}_2\text{O}_5 + \text{I}_2\text{O}_5 = \text{I}_2 + 5\text{N}_2\text{O}_4$, or (II) $10\text{NO} + 3\text{I}_2\text{O}_5 = 3\text{I}_2 + 5\text{N}_2\text{O}_4$. No decisive evidence could be obtained to indicate that it occurred according to I. The assumption that first II takes place with the formation of N_2O_4 , and subsequently N_2O_4 is formed according to $5\text{N}_2\text{O}_4 + 5\text{NO} = 7.5\text{N}_2\text{O}_5$ is equally possible. J. B.

Action of mercuric oxide on solutions of cupric chloride. H. PELANON AND M. DELWAULLE *Bull. soc. chim.* [4], 47, 556-9(1930) cf. *C. A.* 19, 1105, 24, 3454 — The reaction between HgO and CuCl_2 is shown to take place according to $3\text{HgO} + 4\text{CuCl}_2 = 3\text{HgCl}_2 + 3\text{CuO} + \text{CuCl}_2$. The action of CuO on, as well as the soly. of HgCl_2 in, solns. of CuCl_2 was studied separately. G. TOERNVIST

Formation of a ferromagnetic iron sesquioxide by decomposition of the hydrate of van Bemmelen's sesquioxide. G. CHAUDRON AND A. GIRARD *Compt. rend.* 192, 97-9 (1931) — The cryst. $\text{Fe}_2\text{O}_3 \cdot \text{H}_2\text{O}$ prepd. from $\text{Fe}_2\text{O}_3 \cdot \text{Na}_2\text{O}$ by the method of van Bemmelen (*J. prakt. Chem.* 46, 520(1892)) loses water slowly at 100° and rapidly at 140° to form a strongly magnetic sesquioxide which at 675° is converted into the ordinary sesquioxide. Both the monohydrate and the oxide are hexagonal, while the magnetic oxide obtained by slow oxidation of magnetite (Malaguti's oxide) is cubical (cf. Welo and Naudisch, *C. A.* 19, 3182).

K. V. THIMANN

Decomposition of magnesium salts by calcium carbonate at elevated temperatures. E. SAUER AND J. HUTER *Z. anorg. allgem. Chem.* 195, 211-6(1931) — MgCl_2 , MgSO_4 , and $\text{Ca}(\text{HCO}_3)_2$ in aq. soln. were autoclaved for long periods (as high as 9 hrs.); the courses of the decomps. were followed. Mixts. of the Mg salts with $\text{Ca}(\text{HCO}_3)_2$ were also studied, pptn. of $\text{Mg}(\text{OH})_2$ occurs. Solid CaCO_3 was also present in some expts. Data and curves are given.

WILLIAM E. VAUGHAN

The double decomposition between aqueous sodium formate and ammonium sulfate. G. BARDIO AND P. LOTZ *Chem.-Ztg.* 55, 4(1931) — Isotherms are given for the reciprocal salt pair at 15° , 21.0° and 60° . There are several difficulties in the way of prepg. HCO_2NH_4 by double decompn.

T. H. CHILTON

Thoryl nitrate. P. CHAUVENET AND MME. SOUTEVRAUD-FRANCK *Bull. soc. chim.* [4], 47, 1129-31(1930) — $\text{Th}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$ was obtained by slow evapn. of a soln. in air at 15° . Above 80° the trihydrate was obtained; above 110° the dihydrate; from 125° to 170° decompn. of the salt yielded $\text{Th}(\text{NO}_3)_3 \cdot 1/2\text{H}_2\text{O}$; from 150° to 300° decompn. was continuous, yielding pure ThO_2 . The dihydrate was not affected by reduced pressure. Anhyd. $\text{Th}(\text{NO}_3)_3$ could not be obtained. $\text{Th}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$ reacted in soln. with 2 mols. NaOH , yielding $\text{ThO}(\text{NO}_3) \cdot \text{H}_2\text{O}$. No other product of hydrolysis was obtained. The hydrolysis is similar to that of ZrCl_4 (cf. *C. A.* 24, 5654).

GERAID M. PRATT

The complex cuprothiosulfates of ammonium, potassium and sodium. III. G. SPACU AND I. G. MURGULESCU *Bul. soc. stiinta Cluj* 5, 311-70(1930); cf. *C. A.* 24, 4230 — Potentiometric titrations of $\text{Cu}(\text{NO}_3)_2$ by $\text{Na}_2\text{S}_2\text{O}_3$ in the presence of regulated amts. of EtOH show flexures corresponding to the following Na cuprothiosulfates: $\text{Na}_2\text{Cu}_2(\text{S}_2\text{O}_3)_2 \cdot 8\text{H}_2\text{O}$, $\text{NaCuS}_2\text{O}_3 \cdot 1/2\text{H}_2\text{O}$, $\text{Na}_2\text{Cu}_2(\text{S}_2\text{O}_3)_2 \cdot 11\text{H}_2\text{O}$, $\text{Na}_4\text{Cu}_2(\text{S}_2\text{O}_3)_2 \cdot 11\text{H}_2\text{O}$, $\text{Na}_4\text{Cu}_2(\text{S}_2\text{O}_3)_2 \cdot 2\text{NaNO}_3$. $\text{Na}_2\text{Cu}_2(\text{S}_2\text{O}_3)_2 \cdot 8\text{H}_2\text{O}$ is formed as a yellow cryst. ppt. when 7 vols. of 0.5 M $\text{Cu}(\text{NO}_3)_2$ and 13 vols. of 0.5 M $\text{Na}_2\text{S}_2\text{O}_3$ are mixed in the presence of 3 1/2 vols. of EtOH . $\text{NaCuS}_2\text{O}_3 \cdot 1/2\text{H}_2\text{O}$ is obtained as a white amorphous ppt. when 5 vols. of 0.2 M $\text{Cu}(\text{NO}_3)_2$ are mixed with 4 vols. of 0.5 M $\text{Na}_2\text{S}_2\text{O}_3$ in the presence of 16-20 vols. of EtOH . When 2 moles of $\text{NaCuS}_2\text{O}_3 \cdot 1/2\text{H}_2\text{O}$ are titrated with 1 mole of $\text{Na}_2\text{S}_2\text{O}_3$, $\text{Na}_4\text{Cu}_2(\text{S}_2\text{O}_3)_2$ seems to be produced in the form of a brown oily soln. This salt could not be further isolated, for upon evapn. the soln. changed to a glassy homo-

geneous mass. It was attempted to prep $\text{Na}_2\text{Cu}(\text{S}_2\text{O}_3)_2$ from a soln. contg theoretical amts of $\text{Na}_2\text{S}_2\text{O}_3$ and NaCuS_2O_3 . A mass of iridescent crystals was obtained, too hygroscopic for further purification. The double salt $\text{Na}_2\text{Cu}(\text{S}_2\text{O}_3)_2 \cdot 2\text{NaNO}_3$ is prepd by treating a mixt of NaCuS_2O_3 and NaNO_3 with 3 M $\text{Na}_2\text{S}_2\text{O}_3$. Upon recrystn it appeared as colorless needles. The following table contrasts the hydration of the cuprothiosulfates of NH_4 , K and Na.

$\text{NH}_4\text{CuS}_2\text{O}_3 \cdot \frac{1}{2}\text{H}_2\text{O}$	$\text{KCuS}_2\text{O}_3 \cdot \text{H}_2\text{O}$	$\text{Na}_2\text{Cu}_2(\text{S}_2\text{O}_3)_3 \cdot 8\frac{1}{2}\text{H}_2\text{O}$
$(\text{NH}_4)_2\text{Cu}_2(\text{S}_2\text{O}_3)_3 \cdot 2\text{H}_2\text{O}$	$\text{K}_2\text{Cu}_2(\text{S}_2\text{O}_3)_3 \cdot 2\text{H}_2\text{O}$	$\text{NaCuS}_2\text{O}_3 \cdot 1\frac{1}{2}\text{H}_2\text{O}$
	$\text{K}_2\text{Cu}_2(\text{S}_2\text{O}_3)_3 \cdot \text{H}_2\text{O}$	$\text{Na}_2\text{Cu}_2(\text{S}_2\text{O}_3)_3 \cdot \text{xH}_2\text{O}$
	$\text{K}_2\text{Cu}(\text{S}_2\text{O}_3)_3 \cdot 3\text{H}_2\text{O}$	$\text{Na}_2\text{Cu}(\text{S}_2\text{O}_3)_3 \cdot \text{xH}_2\text{O}$

J. H. REEDY

A series of new tetrammine-cobaltic complexes. PULIN BHABH SARKAR AND TUKHANGSU DAS GUPTA *J Indian Chem Soc* 7, 835-7(1930) — $\left[\text{Co}^{\text{Cl}}(\text{NH}_3)_4\right]\text{S}_2\text{O}_8$

at 35-40° for 4 weeks changes completely to the more stable $\left[\text{Co}^{\text{S}_2\text{O}_8}(\text{NH}_3)_4\right]\text{Cl}$ without appreciable change in wt. Thiosulfato-pentammine compds possess a characteristic ultra violet absorption spectrum by which they can be easily distinguished from their ionic isomers. Their conductivities are half that of the corresponding chloro compd

$\left[\text{Co}^{\text{Cl}}(\text{NH}_3)_5\right]\text{Cl}$ or $\left[\text{Co}^{\text{Cl}}(\text{NH}_3)_4\right]\text{Cl}_2$ in aq soln with $\text{Na}_2\text{S}_2\text{O}_3$ below 0° gives trans

thiosulfato-aquo tetrammine cobaltic thiosulfate $\left[\text{Co}^{\text{S}_2\text{O}_8}(\text{NH}_3)_4\right]\text{S}_2\text{O}_8$ (I) by double

decompn. The free S_2O_8 group of I is replaced to give the corresponding Cl, Br, NO₃, SCN, CrO₄, SO₄ and S₂O₃ compds. Their ultra violet absorption spectra in aq solns are identical. At higher temps the soln by hydrolysis gives

$\left[\text{Co}^{\text{S}_2\text{O}_8}(\text{NH}_3)_4\right]\text{OH}$. Further hydrolysis permits the formation of H_2SO_4 which partially reduces the Co^{+++} compds to the Co^{++} - Co^{+++} complex, $\left[\text{Co}^{\text{S}_2\text{O}_8}(\text{NH}_3)_4\right]\text{OH}$. $\left[\text{Co}^{\text{S}_2\text{O}_8}(\text{NH}_3)_4\right]\text{OH}$ isolated by a/c

FOSTER DEE SNELL

A new class of amines. The double selenocyanammines. G. SPACU AND V. ARMEANU *Bul soc chimie Cluj* 5, 294-318(1930) — Solns of $\text{K}_2\text{Hg}(\text{SeCN})_4$ and $\text{KAg}(\text{SeCN})_2$ were formed by the action of KSeCN on HgCl_2 and AgCl , resp. By causing these solns to react with various metallic salts of ethylenediamine, the following double selenocyanammines were pptd: $\text{Cu}(\text{en})_2[\text{Hg}(\text{SeCN})_4]$, $\text{Ni}(\text{en})_2[\text{Hg}(\text{SeCN})_4]$, $\text{Cd}(\text{en})_2[\text{Hg}(\text{SeCN})_4]$, $\text{Zn}(\text{en})_2[\text{Hg}(\text{SeCN})_4]$, $\text{Co}(\text{en})_2[\text{Hg}(\text{SeCN})_4]$, $\text{Cu}(\text{en})_2[\text{Ag}(\text{SeCN})_2]$, $\text{Ni}(\text{en})_2[\text{Ag}(\text{SeCN})_2]$, $\text{Cd}(\text{en})_2[\text{Ag}(\text{SeCN})_2]$, $\text{Zn}(\text{en})_2[\text{Ag}(\text{SeCN})_2]$, $\text{Co}(\text{en})_2[\text{Ag}(\text{SeCN})_2]$. These double selenocyanammines as a class, are cryst solids, stable toward light and dry air. They are insol in EtOH, Et₂O, C₆H₆, CHCl₃, etc., and slightly sol in H₂O, Me₂CO, etc. They are decompd by the feeblest acids and bases.

J. H. REEDY

The salts of hexafluorophosphoric acid. WILLI LANGE AND EMIL MÜLLER *Ber* 63B, 1058-70(1930) — The reaction of chl HF with P_2O_5 gives three types of acids $\text{H}_2\text{PO}_3\text{F}$, HPO_2F_2 and HPF_4 , $\text{H}_2\text{PO}_3\text{F}$ and HPF_4 ppt with nitron. Keeping the soln overnight results in complete hydrolysis of HPO_2F_2 , and the soln yields with nitron a ppt of pure nitron HPF_4 . This ppt shaken with chloroform and ammonia gives the well crystallizing NH_4PF_6 in the water soln. The hexafluorophosphates strikingly resemble the perchlorates in soly. They differ completely, however, in crystal form. All the inorg., especially the complex, PF_6 compds have refractive indexes similar to that of water, which gives them a gel-like appearance in water suspension. Aryl diazonium PF_6 compds are less sol than the perchlorates, they do not decompose when dry, and are not sensitive to shock or blows. Heating results in decompn according to $\text{ArN}_2\text{PF}_6 = \text{ArF} + \text{PF}_5 + \text{N}_2$. HPF_6 is stable in cold soln but is hydrolyzed in hot soln. PF_6 compds on boiling with strong acids are hydrolyzed. Mol vols of PF_6 and ClO_4 compds are compared. Prepn and properties are described for KPF_6 , NH_4PF_6 , NaPF_6 , H_2O , RbPF_6 , CsPF_6 , $\text{Cu}(\text{PF}_6)_2 \cdot 5\text{NH}_3$, H_2O , $\text{Cu}(\text{PF}_6)_2 \cdot 4\text{NH}_3$, $\text{Cu}(\text{PF}_6)_2 \cdot 2\text{C}_6\text{H}_5(\text{NH}_2)$, $2\text{H}_2\text{O}$, $\text{Cu}(\text{PF}_6)_2 \cdot 5\text{C}_6\text{H}_5\text{N}$, NH_3 , H_2O , $\text{Co}(\text{PF}_6)_2 \cdot 6\text{NH}_3$, $\text{Ni}(\text{PF}_6)_2 \cdot 6\text{NH}_3$, Me_4NPF_6 , Et_4NPF_6 , Ph_4NPF_6 , $\text{p-MeC}_6\text{H}_4\text{NPF}_6$, $\text{p-p'F}_4\text{PN}$, $\text{C}_6\text{H}_5\text{N}$, $\text{C}_6\text{H}_5\text{N}$, NPF_6 , strychnine HPF_6 , brucine HPF_6 , morphine-HPF_6 , cocaine-HPF_6 , nitron HPF_6 , nitron HAsF_6 , H_2O , nitron HSBF_6 , $2\text{H}_2\text{O}$ and nitron HSBF_6 . R. W. S.

Bis-*p*-phenetyl telluride and its derivatives. GILBERT T MORGAN AND FRANCIS H BURSTALL. *J Chem Soc.* 1930, 2599-601 —Bis *p*-phenetyl telluride and some of its derivs are prepd., the view of Burstall and Sugden (*C A* 24, 2652) that bis-*p*-phenetyl telluridichloride is a nuclear-substituted telluride being confirmed. *Bis p phenetyl telluride*, $(\text{EtO C}_6\text{H}_4)_2\text{Te}$, (white plates, m 64°) is prepd by adding 63 g *p*-bromophenetole (by the Sandmeyer reaction from *p*-phenetidine) to Mg (7.2 g) suspended in dry Et_2O , warming until nearly all the Mg is in soln., decanting, adding the soln. slowly to a TeBr_4 (43 g) suspension in dry Et_2O , refluxing for several hrs., cooling, treating with ice, adding HBr , filtering, drying the Et_2O layer over CaCl_2 , evapg to dryness, extg the residue of gum with petroleum ether (b $40-60^\circ$), evapg and repeatedly recrystg the solid sepg mut from alc. *Bis-*p*-phenetyl telluridichloride*, $(\text{EtO C}_6\text{H}_4)_2\text{Te}$, m 108° , is prepd (1) by adding Cl_2 to a soln of the preceding telluride in CCl_4 , (2) by heating *p*-phenetyl telluridichloride with an excess of phenetole at $180-190^\circ$ and (3) by heating TeCl_4 (1 mol) with phenetole (4 mols) at $180-190^\circ$.

Bis p phenetyl telluriodichloride, $(\text{EtO C}_6\text{H}_4)_2\text{Te} \begin{array}{c} \diagup \text{O} \diagdown \\ \text{Cl} \quad \text{Cl} \end{array} \text{Te}(\text{C}_6\text{H}_4 \text{OEt})_2$, (fluffy white

crystals, decompg at 193°) is prepd by boiling an aq suspension of the dichloride. *Bis p-phenetyl telluridibromide*, $(\text{EtO C}_6\text{H}_4)_2\text{Br}_2\text{Te}$, (lemon yellow crystals, m 127°) and *bis p-phenetyl telluridiiodide*, $(\text{EtO C}_6\text{H}_4)_2\text{I}_2\text{Te}$, (orange red needles, m 144°) are prepd from mixts of their generators in soln. *Bis p-phenetyl telluroxide*, $(\text{EtO C}_6\text{H}_4)_2\text{TeO}$, (white microcrystals, decompg at 181°) is prepd by boiling the dichloride or the dibromide with aq 2*N* NaOH soln. *p-Anisyl p-phenetyl telluridichloride* (white needles, m $165-6^\circ$) is prepd by heating the trichloride (1 mol) at 160° with anisole (4 mols) for 6 hrs., cooling, treating with petroleum ether (b $40-60^\circ$) and recrystg the residue repeatedly from alc. *p-Anisyl p-phenetyl telluride* (white needles, melting to a clear liquid at 45°) is prepd by heating the dichloride for an hr in glacial AcOH with Zn dust in excess, filtering, pouring the filtrate into H_2O and crystg the ppt from alc.

J. BALOZIAN

Researches on residual affinity and coördination. XXXII. Complex salts of bivalent silver. GILBERT T MORGAN AND FRANCIS H BURSTALL. *J Chem Soc* 1930, 2594-8, cf. *C A* 23, 1877 —As with Cu salts and a diamine (*C A* 20, 3400, 3401, 21, 3166), it has been found that the chelate group in α, α' -dipyridyl is sufficiently powerful to stabilize a bivalent argentic complex. Thus, by starting with α, α' -dipyridyl (as prepd by Hein and Retter (*C A* 23, 143)), a series of Ag salts is prepd in which the bivalent Ag has coördination nos of both 4 and 6. *Bis- α, α' -dipyridylargenticous nitrate* $[\text{Ag}_2\text{dipy}]\text{INO}_3$, (felted masses of yellow needles, decompg at 153°) is prepd from hot alc. solns of AgNO_3 (16.9 g) and α, α' -dipyridyl (31.2 g), and recrystd from hot dil alc. *Bis α, α' -dipyridylargentic persulfate*, $[\text{Ag}_2\text{dipy}]\text{S}_2\text{O}_8$, (deep reddish brown microcrystals, decompg. rapidly at 137°) is prepd by stirring the preceding nitrate into a cold satd soln. contg. an excess of $\text{K}_2\text{S}_2\text{O}_8$. *Pentakis- α, α' -dipyridylargentic persulfate*, $[\text{Ag}_5\text{dipy}](\text{S}_2\text{O}_8)_3$, (small reddish purple needles, decompg rapidly at 159°) is prepd by triturating the preceding persulfate with cold dil. HNO_3 (60% HNO_3 , $d = 1.4$) and 40% H_2O) and pptg. with H_2O (4 vols). *Tris- α, α' -dipyridylargentic nitrate*, $[\text{Ag}_3\text{dipy}]\text{H}(\text{NO}_3)_2$, (small dark brown needles, decompg at 176°) is obtained by draining the residue from the persulfate trituration, extg with warm H_2O , adding an excess of aq NH_4NO_3 and cooling in ice. *Tris- α, α' -dipyridylargentic chloride*, $[\text{Ag}_3\text{dipy}]\text{Cl}_3$, (dark purple plates or needles, decompg quietly at 171°) is prepd by adding aq NaClO_2 to the preceding nitrate and recrystg the product from H_2O by the addn of NaClO_2 . *Tris- α, α' -dipyridylargentic perchlorate*, $[\text{Ag}_3\text{dipy}]\text{H}(\text{ClO}_4)_2$, (orange-brown microcrystals, detonating violently on heating) is prepd by adding NaClO_4 to a soln of either the preceding chlorate or nitrate. *Bis α, α' -dipyridylargentic hydrogen sulfate*, $[\text{Ag}_2\text{dipy}]\text{H}(\text{HSO}_4)_2$, (brown cryst. powder, decompg at 157°) is prepd. by treating the persulfate with cold 40% H_2SO_4 or by adding Na_2SO_4 and H_2SO_4 to solns of the preceding chlorate or nitrate.

J BALOZIAN

7—ANALYTICAL CHEMISTRY

W. T. HALL

Reply to the remarks of L. Moser and L. v. Zombory concerning my paper on "New ways in analytical chemistry." J. DICK. *Z anal. Chem* 83, 105-7(1931), cf. *C A* 25, 891; M and Z, *C A* 24, 4725 —The method of washing the ppt. with alc. and ether with vacuum drying at room temp is defended

W. T. H

Collecting average samples from storage tanks. H. O. BERNSTROM. *Chemist Analyst* 20, No. 2, 10-1(1931).—The taking of a sample from tank cars by means of a circular thief is described. W. T. H.

Standardizing silver nitrate solutions. ALVIN MAYNAN. *Chemist Analyst* 20, No. 2 14 5(1931).—One soln. of AgNO_3 can be standardized against another of the same kind by making an approx. 0.21 N HCl soln. and titrating one portion with the old and another portion with the new soln. by Mohr's method. W. T. H.

Sodium sulfite for analysis. HOLZWECKHOLMUNG-IND A-G. *Chem.-Ztg.* 54, 502(1930).—Contrary to earlier statements, pure Na_2SO_3 does not yield alk. solns., and the faint blue color with thymolphthalein is destroyed by one drop of 0.1 N HCl . In the use of Na_2SO_3 for the detn. of CH_2O , the grade "pro analysi" is not essential, because reasonable proportions of neutral saline impurities, such as Na_2SO_4 , do not affect the titration, Na_2CO_3 must, however, be absent. B. C. A.

"Direct green B." A new reagent very sensitive to copper. P. SISLEY AND DAVID. *Bull. soc. chim.* [4], 47, 118-92(1930).—When Cu salts are added at 90° to a soln. of "direct green B," the color changes from blue-green to violet pink. The soln. should have a pH value between 6 and 9. The ratio dye/ Cu should be 10, but results are satisfactory between 6 and 50. The reaction is sensitive to 0.1 mg. Cu per l. Anions had no effect on the color. No other metal gave a violet pink color. Neutral salts retarded the reaction and reduced the intensity of the color, K_2SO_4 having the greatest effect. Colloids masked the reaction. Cu was found in distd. water used for diluting serums and medicines. GERALD M. PETTY.

Correct methods of analyzing for carbon in rustless steels. C. M. JOHNSON. *Iron Age* 127, 549-51(1931).—In some cases it is necessary to have the results on C detns. accurate to 0.01% of the wt. of the sample. Low results may be caused by incomplete combustion because of failure to cover the sample with PbO_2 , and high results may be caused by the S content. Directions are given for carrying out the combustion in an elec. furnace, and an absorption train is shown in which danger due to SO_2 or SO_3 fumes is overcome by passing the gas through 50 ml. of 5% KMnO_4 soln., through the usual Zn tower, through P_2O_5 , and finally into the Stetser and Norton ascarite bulb. W. T. H.

Determination of total sulfur by the sodium peroxide method. M. R. BERCOVITCH. *Chemist Analyst* 20, No. 2, 8(1931).—In neutralizing the aq. ext. of the melt, loss by effervescence can be avoided by introducing the acid at the bottom of the soln. by means of a pipet or funnel drawn out to a fine point. W. T. H.

Wet oxidation method for the determination of total sulfur in slags, ores, cinders, etc. FRANK W. SCOTT. *Chemist Analyst* 20, No. 2, 8(1931).—The sample is dissolved in concd. HNO_3 which has been saturated with KClO_3 , and the resulting SO_4^{--} is detd. as in the well known Lunge method. W. T. H.

Detection of minute quantities of phosphorus in oil. CONRAD STICH. *Pharm. Ztg.* 76, 112(1931).—A commentary on S.'s previously published method (*C. A.* 21, 3856) with suggestions for its more precise application. W. O. F.

Spectrophotometric determination of phosphorus. TORSTEN TEORRELL. *Biochem. Z.* 230, 1-9(1931).—Fiske-Subbarow's procedure for the colorimetric detn. of P is modified to be used for a photospectrometric analysis of quantities ranging from 0.01 to 0.05 mg. P . The method dispenses entirely with the standard soln., and detns. can be made with an accuracy of $\pm 2\%$. S. MOAGULIS.

Colorimetric estimation of nitrogen by direct nesslerization, with a note on a modified Nessler-Folin reagent. LANCE S. WALTERS. *Australian J. Exptl. Biol. Med. Sci.* 7, 113-6(1931).—By using a gum ghatti soln. for protective colloid and replacing NaOH by LiOH in the Nessler-Folin soln., it was possible to nesslerize 100 cc. of soln. contg. 2 cc. concd. H_2SO_4 , and up to 1 mg. N . The nesslerized soln. remained clear for weeks. The usual Kjeldahl digest contg. 20 cc. of H_2SO_4 is diluted to 100 cc. and an aliquot taken which contains 0.5 to 1.0 mg. N . This is diluted to about 60 cc., 3 cc. colloid soln. is added, followed by nesslerization and making up to 100 cc. for matching a standard. The protective colloid soln. contains 2% of alc.-pptd. and dried gum in thymol-satd. water. C. G. KING.

Determination of nitrogen by a "micro-Kjeldahl" method. A. C. ANDERSEN AND B. NORMAN JENSEN. *Z. anal. Chem.* 83, 114-20(1931).—The method previously described by A. and J. (*C. A.* 19, 1461) can be used equally well for 0.5-1 mg. of N if the app. is built on a smaller scale. Such a modification is shown and exact dimensions are given. The app. was tested with 5 org. nitrogenous compds. and excellent results were obtained. W. T. H.

Determination of some gases by means of the spectrograph. BOUCHETAL DE LA

ROCHE *Bull. soc. chim.* [4], 47, 1326-31 (1930), cf *C. A.* 24, 1252, 1580, 5667.—Further description of the method is given. Spectrographic analyses of mixts of Si and B fluorides and Cl with dry air are in close agreement with the results by the chem. method. Br and I are less sensitive and therefore difficult to det. by this method. O gives certain lines by which it can be detd., but the intensity is independent of the quantity of O introduced, this is probably due to the oxidation of the electrodes. N diluted with H₂ or CO₂ is detd. by the line 5008 or 3437, resp., and the results agree well with those of volumetric analysis.

I. W. WALKER

Bromodimetric investigations. V. Determination of bromine in bromides. J. H. VAN DER MEULEN *Chem. Weekblad* 28, 82-6 (1931), cf *C. A.* 25, 1182.—To 20 cc. KBr soln (contg. about 0.1 g.) add 10 cc. N KOCl in 0.1 N KOH and 25 cc. satd NaCl soln. Heat the mixt. on the water bath to about 85° and add 10 cc. of a satd H₂BO₃ soln. Continue the heating for about 5 min. after discoloration takes place. Decompose the excess KOCl with 10 cc. N H₂O₂ and boil off the excess H₂O₂. After cooling add KI and HCl and titrate the mixt. A blank titration is necessary.

F. SCHIOTTE

Spectroscopic detection of fluorine. JACOB PAPISH, L. E. HOAG and W. E. SWEET, *Ind. Eng. Chem., Anal. Ed.* 2, 263-4 (1930).—A mixt. of a substance contg. F with a salt of Ca is placed on a graphite electrode and the arc or spark spectrum photographed. A band at 5291 Å. U. is intense enough to detect 0.02 mg. F. The method is applied to complex F compds., F minerals, bones and teeth of animals. In the latter case and when the quantity of F is very small, the sample is concd. by distg. with H₂SO₄ and the HF is collected in Ca(OH)₂ soln. The dried CaF₂ is then applied to the electrodes.

G. M. MURPHY

Qualitative microanalysis. Detection of traces of copper by capillary separation. F. FEIGL and H. J. KAPULITZAS *Mikrochemie* [N. S.], 2, 239-44 (1930).—Traces of Cu may be detected in the presence of large amts. of Ni by placing one drop of the AcOH soln. of the elements on a filter paper which has been dipped into an aq. soln. of dithiodiamide and dried. If Cu is present, the middle of the drop immediately develops an olive green to black ring, the Ni diffuses further before it reacts with the amide, so that a white ring surrounds the Cu ring and is itself surrounded by a dirty violet to violet blue ring. The reaction will detect 0.03% of Cu in the presence of 20,000 times as much Ni or 2000 times as much Co, which forms a yellowish brown ring outside the Cu ring.

B. C. A.

Effect of the presence of other elements on some microscopical tests for the metals. W. F. WHITMORE and F. SCHNEIDER *Mikrochemie* [N. S.], 2, 293-304 (1930).—The formation of triple nitrates with K and Pb is a test for Cu or Ni in the absence of the other, the former giving black or dark brown cubes and the latter light yellow cubes; when both elements are present the color may vary from light yellow to black. The oxalate test for Mn, Cd, Pb and bivalent Sn is sp. only when each element is present alone; when 2 or more are present in the test soln. the crystals formed usually have a different habit from the normal, or the crystals due to one element may fail to appear. Similarly, the Hg(CNS)₂ test for Cu, Co, Zn and Cd does not distinguish the elements from one another when in admixt., the crystal habit and color varying considerably with the nature and relative proportions of the metals present. Fe cannot be detected by the ferrocyanide test when more than 10 times as much Cu is present, and Pb and Hg cannot be detected by the iodide test when 50 times as much Hg or Pb is present in the same soln. (cf. *C. A.* 24, 2964).

B. C. A.

Analysis of white metals and solder. H. L. MATTHIJSEN *Chem. Weekblad* 27, 281-7 (1930).—Slight modifications in standard methods, and particularly in simultaneous electrolytic detn. of Pb and Cu, are described.

B. C. A.

Spot tests for the detection of the precious metals. H. HOLZER *Mikrochemie* [N. S.], 2, 271-6 (1930).—For the detection of Au in a soln. contg. Pt a drop of the liquid is placed on a filter paper and allowed to dry; on addn. of a min. drop of acid SnCl₄ soln. to the middle of the spot it becomes yellowish brown, and a purple ring gradually forms around the outside if Au is present. Feigl's test for Ag with *p*-dimethylamino-benzylideneazobenzene (C. A. 22, 4080) is interfered with by Au, Pt and Pd, which produce colored flocculent ppts. with the reagent. The Pd ppt. is dark violet in color and its formation serves as a sp. test for Pd.

B. C. A.

Inorganic gravimetric microanalysis. I. Determination of minute amounts of gold in the presence of much iron, lead and copper. J. DONAU *Mikrochemie* [N. S.], 2, 257-63 (1930).—To the boiling, slightly acid soln. (0.2-0.3 cc.) of the Au is added a small piece of Japanese silk previously washed in HCl, then in water and dried. The Au deposits rapidly on the silk, coloring it dark purple; a 2nd piece of silk may be added

to make sure that pptn is complete. The alk is collected on a glass microfilter, washed with hot water, dried in a small Pt foil dish and ignited over a small flame. Further heating causes the Au residue to alloy with the Pt, which is then weighed on the microbalance. Fe, Cu and Pb do not interfere in the test. B. C. A.

Determination of cadmium with potassium ferrocyanide. ALFRED KUNDERT. *Chemist Analyst* 20, No. 2, 5-6(1931)—It is shown that the $K_4Fe(CN)_6$ titration gives results that are comparable to those obtained gravimetrically. W. T. H.

Specific reaction for the detection of cadmium. F. PAVELKA AND E. KOLMER. *Mikrochemie* [N. S.], 2, 277-80(1930)—Cd salts in AcOH soln. yield a characteristic yellowish white cryst. ppt. with nitrophenolarsinic acid, under the microscope the crystals appear as long narrow rectangular prisms often clustered in prickly spherical aggregates. The reaction will detect 1 part of Cd in 200,000 parts of soln. and is not affected by the presence of any metal which forms a sol acetate. B. C. A.

Detection of zinc by means of resorcinol. K. FAHREN. *Verhandl. Geol. Bundesanstalt Wien* 1929.—The addn of a 10% soln of resorcinol in alc. to a dil. soln. of a Zn salt gives on shaking in the cold a yellow color, which on heating changes to olive green or blue, the color depending on the concn. Other metals, such as Cd, Ni, Co, Cu and Mn, give similar colors with resorcinol. B. C. A.

Analysis of the coating of galvanized sheet steel. FRANK W. SCOTT. *Chemist Analyst* 20, No. 2, 4-5(1931)—Cut 3 test pieces, 2.25 in. sq., of the metal, clean with gasoline or benzene, dry, mark one of them and weigh. Place the 3 pieces in a 400-cc. beaker with their bottom edges resting on pieces of pure Zn and with the sheets sep'd. by glass rods. Cover the squares of metal with 3.3 N H_2SO_4 and put a watch glass over the beaker. Move the pieces from time to time but keep them in close contact with the pure Zn. When the evolution of gas is very slow, showing that the coating has been dissolved, transfer the acid to another beaker and at once cover the undissolved metal with water. Using a policeman and a jet from the washer bottle, scrub off the sheets and pieces of pure Zn, adding the black residue and water to the acid soln. Dry the marked sheet and again weigh, the loss in wt. in g. is equal to the wt. of coating in oz. on a piece a foot sq. This wt. multiplied by 3 gives the wt. of sample analyzed. Titrate the Fe^{++} in the soln with $KMnO_4$. Add HNO_3 , evap with conc'd H_2SO_4 , and det the Pb as $PbSO_4$. Dil the filtrate from the $PbSO_4$ to 450 cc and, while introducing H_2S , reduce the acidity so that some Zn together with all the Cd is pptd as sulfide. Filter, dissolve in acid and repeat this pptn twice more until finally pure CdS is obtained and weighed. To det. Sn, treat 3 more pieces in the same way, from the resulting soln. ppt. sulfides of Sn and Pb with H_2S , dissolve the ppt. in aqua regia, boil off Cl_2 , dil., nearly neutralize with NH_4OH , add an excess of NH_4OAc , make neutral to methyl orange and allow to stand in a warm place to ppt. $Sn(OH)_2$. Filter, ignite and weigh as SnO_2 . Call the remainder of the coating Zn. W. T. H.

Determination of the iron contained in the coating of galvanized steel. BENJAMIN E. COUV. *Chemist Analyst* 20, No. 2, 4(1931)—The method of obtaining the sample is similar to that described by Scott (see preceding abstr.) but $HgCl_2$ is added to the soln. before titrating with $KMnO_4$. W. T. H.

The determination of iron in aluminum. GIULIO AGAMENNONE. *Giorn. chim. ind. applicata* 12, 486-7(1930)—Fe occurring in Al is usually det'd volumetrically as Fe^{++} , leading to some losses due to carbides not completely decomposed because of the danger of oxidizing to Fe^{+++} again, and to some Fe remaining as Fe^{+++} . To avoid these losses, Fe is det'd as Fe^{+++} by dissolving the Al in dil. HCl, boiling the soln., adding H_2O_2 and again boiling. The soln. is then cooled and the Fe^{+++} det'd iodometrically. A. W. CONTIERI.

New test for tin. LAWREN STONE. *Chemist Analyst* 20, No. 2, 6-7(1931)—Meissner's Bunsen flame test is described in which a blue coloration is obtained. W. T. H.

Lead-antimony-tin alloys. MAX HALE. *Chemist Analyst* 20, No. 2, 10(1931)—The sp. gr. of the sample gives a fair idea of the Sb content of Pb-Sb alloys. The treatment of the alloy with acid, sep'n of Sb and Sn from Pb by NaOH soln., detn. of Pb as $PbSO_4$ in the ppt., electrodeposition of the Sb from the alk. soln. to which KCN is added and the pptn. of Sn in the electrolyzed soln. are described. W. T. H.

Electrolytic determination of copper in steel. LAWRENCE ANDERSON. *Chemist Analyst* 20, No. 2, 7-8(1931)—Dissolve 5 g. of metal in 100 ml. of 3 N H_2SO_4 , filter off the Cu and any insol. carbide, treat the residue with hot dil. HNO_3 , filter and electrolyze. W. T. H.

Analytical applications of a reaction of copper salts. MANUEL LORA Y TAMAYO. *Rev. acad. ciencias Madrid* [2], 25, 267-98(1930)—The reaction between Cu^{++} and

tartrate in alk. soln serves to indicate 15.0-16.7 mg of Cu^{++} per 100 cc. of aq. soln, or 6.4 mg. of Cu^{++} if sufficient alk. is added. By measuring the vol. of KCN used to decolorize the soln a fairly satisfactory detn. of Cu^{++} can be made. A luidiography of methods for detg. Cu and of studies of the action of various cations on alk. tartrate is given.

JOHN H. FRIEDRICH

Study of Spacu's reaction. Methods for the volumetric determination of copper based on this test. J. GOSN. *Bull. soc. chim.* 40, 81-100 (1931).—Salts of Cu give ppt. of $[\text{Cu}(\text{C}_6\text{H}_5\text{N})_4](\text{SCN})_2$ when placed in contact with pyridine and an alk. thiocyanate. The ppt. is easily dissolved by CHCl_3 giving an emerald green soln., and by evapn. of this soln. a residue is obtained which can be weighed as such or ignited to CuO . Quantities of Cu varying from 5 to 60 mg. can be detd. volumetrically by using 1 cc. of pyridine and 25 cc. of 0.1 N KCN in a total vol. of 100 cc., filtering off the ppt., adding a known quantity of AgNO_3 and titrating the excess Ag^+ with 0.05 N NH_4CNS , using ferric alum as indicator. Cl^- , Br^- , Zn , Cd , Ni and Co must be absent and proper allowances must be made for a blank run in the absence of Cu. Another method, suitable for even smaller quantities of Cu, consists in pptg. the Cu similarly, filtering the soln. and dtg. the excess CNS by the following method: Take 20 cc. of filtrate, add 0.5 cc. of NaOH (d 1.02-1.30) and expel excess pyridine by a current of steam, using glass connections without any contact with rubber and keeping the vol. at 20 cc. Cool and dil. to 30 cc. Add 20 cc. NaBrO_3 soln. (5 cc. Br_2 , 50 cc. of NaOH (d 1.32-1.36) and water to make 500 cc.) to oxidize the excess CNS to SO_4^{--} and CNO^- with reduction of BrO_3^- to Br^- . After 15 min. add 10 cc. of 20% KI and 20 cc. of 2 N HCl. Titrate the liberated I_2 with 0.1 N $\text{Na}_2\text{S}_2\text{O}_3$. Run a blank without any Cu soln.

W. T. H.

Separation of nickel from cobalt metal or ferro-cobalt. H. A. KAR. *Chemist Analyst* 20, No. 2, 15 (1931). Treat the HNO_3 soln. with NH_4OH and $(\text{NH}_4)_2\text{S}_2\text{O}_8$, filter, add considerable NH_4Cl and treat with dimethylglyoxime. Dissolve the ppt. in acid and repeat the above treatment. The final Ni ppt. will be free from Co.

W. T. H.

Estimation of copper and nickel in steel. W. L. KNOX. *Paper before Swansea Tech. Coll. Met. Soc.*, 18 pp. (Feb. 22, 1930), *Metals & Alloys* 1, 734.—If small amounts of Cu are assayed with large units of Fe, the sepn. is best effected by the pptn. of the Cu as a sulfide by means of ZnS in place of the usual NiS sulfocarbonate. No new method is offered for the estn. of the Ni.

G. G.

New volumetric method for the estimation of sodium. A. BLENKINSON. *J. Agr. Sci.* 20, 531-6 (1930).—Certain disadvantages of the gravimetric $\text{NaZn}(\text{UO}_2)_2\text{Ac}$ method (C. J. 22, 2808, 23, 462, 24, 5912), due to manipulative details and interference by a proposed method in which the U is detd. by reduction to U^{++} by means of standard TiCl_3 soln. Ppt. the Na as $(\text{UO}_2)_2\text{ZnNaAc}$, filter, as in the gravimetric method, filter through a Jena glass filtering funnel (porosity G4), remove as much of the reagent as possible by suction and wash 2 or 3 times with 2 cc. portions of a satd. soln. of the ppt. in 95% alc. Dissolve the ppt. in dil. HCl, transfer to a 500 cc. volumetric flask, add about 10 cc. of strong HCl and displace the air in the flask by a stream of CO_2 . Reduce the U to the uranous state by addn. of an excess of 0.02 N TiCl_3 soln. Main-taining the CO_2 atm., shake the mixt. and allow to stand for 2 or 3 min. or heat to 40-50° to insure complete reduction. Add 20-25 cc. of 2% H^+ to stabilize, or even ppt., the U. After addn. of about 10 cc. of 10% KSCN, est. the excess of TiCl_3 by back titration with 0.02 N Fe alum soln. to a permanent red color.

P. R. DAWSON

Determination of water in glycerols. L. F. HOYT and P. C. CLARK. *Soap* 7, No. 2, 23-27, *Oil & Fat Ind.* 8, No. 2, 50-61 (1911).—Water in coned glycerols may be detd. directly by the Bidwell-Sterling distn. method using toluene as the boiling liquid. A variable but small amt. of glycerol (never exceeding 0.1 mg.) is carried over with the water, and the correction of the apparent percentage did not exceed 0.16% in any of the glycerols contg. up to 10% of water. The use of xylene at above 100° is not advisable, as more glycerol is carried over and the time of distn. is not markedly shortened. An analysis of a 25 g. sample of dynamile glycerol to which 2.8 g. of water was added giving 10.37% water, gave a recovery, corrected for 0.030 g. of glycerol carried over, of 10.39% water.

H. SCHERER

Determination of phosphoric acid as ammonium phosphomolybdate or as phosphomolybdic anhydride. A. von KNOX. *Z. anorg. allgem. Chem.* 194, 239-57 (1930).—For pptg. 0.003-0.035 P_2O_5 , it is recommended to add 4 g. NH_4NO_3 , 20 cc. 4 N HNO_3 , and 20 cc. of reagent (100 g. tartaric acid, 150 g. $(\text{NH}_4)_2\text{MoO}_4 \cdot 4\text{H}_2\text{O}$, 10-12 drops HNO_3 per l.) Dil. to 100 cc., shake vigorously and filter after the ppt. has settled. Wash the ppt. with 5% NH_4NO_3 which is 0.0 N in HNO_3 , dry with acetone and weigh.

or ignite. The dried ppt contains 3.55% P_2O_5 , after ignition it contains 3.92% P_2O_5 . The tartaric acid is added to stabilize the soln. B. A. SOULE

Analysis of gases and vapors which affect oxidation of phosphorus. J. TAUSZ AND H. GÖRLACHER. *Z. anal. Chem.* 83, 81-92 (1931).—A method is given for the analysis of substances like hydrocarbons and $Fe(CO)_5$ which tend to prevent oxidation of P in mixts with O_2 and air. As indication of the beginning of its oxidation, the luminescence shown by the P can be taken. The pressure of the gas at which this takes place is sharply defined and reproducible so that it can be called the "luminescence pressure." It serves to detect low concns. of impurities which would be difficult to detect in other ways. The literature on this subject of oxidation hindrance from 1797 to the present time is reviewed. The luminescence pressures of mixts of O_2 and air with the following substances were detd.: benzene, cyclohexane, acetylene, methylcyclohexane, ethylene, propylene, cyclohexene, isoprene and $Fe(CO)_5$. On the basis of these data it is possible to analyze binary gas mixts of these substances with either O_2 or air. The results quoted were accurate within about 0.05%. W. T. H.

Absorption spectrum of sulfohemoglobin. JEAN PRÉVOT. *J. chim. phys.* 54, 576-80 (1930).— H_2S leaves the 2 characteristic bands of hemoglobin unchanged, but adds a new one at 0.610-20 μ , the red blood color being changed to olive-green. In analyzing suspected gas mixts for H_2S the gas was washed with 200 cc. blood soln. in a gas-absorption app. A paraffin oil was used as the replacement fluid, instead of water. A min. of 30 cc. H_2S (room temp. and pressure) produced the characteristic band. CO_2 and O_2 did not interfere. K. H. ENGEL

Notes on the detection of carbon disulfide. F. FRIGL AND K. WEISSELBERG. *Z. anal. Chem.* 83, 93-104 (1931).—The sensitiveness of tests for CS_2 was studied and of 12 methods, the formation of a difficultly sol. Ni salt with "Hector's base," CaH_2N_2S , proved the most sensitive test. In a small test tube of Jena glass, which is provided with a well fitting rubber stopper, put a little of the soln. to be tested, a few crystals of $Ni(OAc)_2$ and some of the org. reagent. If the $Ni(OAc)_2$ is insol. in the liquid, add 5-10 drops of water. Stopper the tube, heat a few min. in a water bath, then remove and allow to cool. A pink color will appear if 0.0005 mg. of CS_2 is present in 1 cc. of soln. tested and a red ppt. if much more is present. The ppt. can be weighed and used as the basis of a gravimetric detn. Of the other methods tested, the best best results were obtained by the reactions with Na_2N_2 and I_2 , in which the CS_2 acts as catalyst, and with $HCHO$ and Na_2HPO_4 ; these reactions served to detect 0.002 mg. of CS_2 in 1 cc. of liquid. W. T. H.

A new reaction to differentiate between kaolin and talcum. E. TACHSCH. *Seifen-Industrie* 57, 909-10 (1930).—Boil about 1 g. of the substance with 10 cc. of 2 N HCl, filter, neutralize with AcOH and add 1 cc. of a freshly prepd. 1% alc. soln. of hematoxylin (ext. of "Blue Wood"), a rose violet color turning blue-violet (test for Al) indicates kaolin, a yellowish color indicates talcum. P. ESCOFFIER

The chemical analysis of bastnasite. I. P. ALTMAN. *Trans. Inst. Econ. Mineral. (Moscow)* No. 44, 55-6 (1930).—P is sep'd from the decompd. mineral according to Berzelius and detd. by the Merwin colorimetric method. Rare earths and sesquioxides are pptd. with NH_4OH and then sep'd by Peter's oxalate method. Th is sep'd. from the Ce and Y groups with H_2O_2 and Ce from the remaining rare earths by Br. Water is detd. by an improved Penfield method. Other common elements are detd. by the usual methods. H. C. PARISH

The determination of sulfate ions in chrome-plating baths. O. MACCINI. *Industria chimica* 5, 1345-51 (1930).—After a lengthy review of the numerous methods which have been proposed for the detn. of SO_4 ions in chrome-plating solns., including an extensive bibliography the following method is developed: To 20 cc. of soln. add 20 cc. of glacial AcOH and reduce Cr(6) with NH_4OH or $CaH_2N_2SO_4$, adding the reagent until there is no further evolution of gas. Add 150 cc. of water, boil and ppt. with 8 cc. of 10% $BaCl_2$. It is claimed that this method will give results within 5% of the truth. A. W. CONTIERI

Analysis by distillation. HANS WOLFF AND J. RABINOWITZ. *Chem. Fabrik* 1931, 37-8.—Directions are given for constructing the simple app. and performing distns. by which the characteristics and boiling relations of solvents, single or mixed, may be detd. when only 2-4 cc. of substance is available. Charts show comparative results by this method and the Engler method for benzene, a mixt. of benzene and acetone, and for a mixt. b. 60-200°, but the nature of this mixt. is not stated. J. H. M.

Determination of the constituents of a mixture of volatile liquids. M. LEMARCHANDS. *Bull. soc. chim.* 49, 80-4 (1931).—The method, which can be used for the detn. of alc. and ether in pastes of Poudre B, depends upon the complete vaporiza-

tion of the alc and ether in a vacuum, the detn of the quantity of heat absorbed by this vaporization and the wt of the substances vaporized. The app consists of a silvered Dewar tube fitted with a stopper carrying a thermometer graduated in tenths of a degree and an Al tube which is provided with an exit tube to be connected with the vacuum pump. The sample is weighed in the Al tube which is placed in the heated Dewar tube which is filled with hot water. The tube is evacuated and the rate of cooling observed. The results obtained were better than those obtained by the usual methods.

W T II

Effect of light on the determination of ethylene. J LOUIS OBERSPIDER AND J. II. BOYD, *Ind Eng Chem, Anal Ed* 3, 123(1931).—By painting the Williams pipet black, except for a narrow vertical strip, the detn can be made more accurate when made in sunlight, which probably increases the activity of Br upon heavy hydrocarbons so that the $C_{12}H_{10}$ results obtained by absorption in Br aq are otherwise high when made in sunlight.

W T II

Volumetric determination of citric acid in calcium citrate. G ROMEO *Rivista ital essenze profumi* 11, xxiii xxvi(1929) cf *C A* 24, 2963. The Ca citrate is converted into the K salt and treated with an excess of standard $AgNO_3$ soln, the excess being titrated with thiocyanate. Oxalic acid, tartaric acid, HCl and H_3PO_4 but not $CaSO_4$, nor pectic substances, interfere.

B C A

Color reaction of glycerol with alkali thiosulfates. D GANASSINI *Arch ist biochim ital* 2, 239-42(1930).—Most of the color reactions of glycerol are due to its products of oxidation or dehydration. Of these, acrolein, which is formed by heating with $KHSO_5$ (or H_2O_2 or I_2O_5) either glycerol or its derivs, may be identified by spontaneous conversion of its phenylhydrazone into phenylpyrazoline this giving an intense violet coloration in the presence of oxidizing agents in acid soln. Powder, $Na_2S_2O_4$ dissolves in very concd glycerol soln, with slight heating to a colorless soln, which becomes sky blue and then deep indigo when heated to boiling. This reaction is not given by carbohydrates in general, but a much less intense coloration is obtained with the polyhydric alcs. To sep glycerol from org liquids and tissues, the material is mixed with water and filtered, the filtrate being treated with a slight excess of basic Pb acetate and filtered. After removal of the Pb as sulfide, the liquid is evapd to dryness with milk of lime, the residue being extd with a mixt of 95% alc (2 parts) and H_2O (1 part) and the soln evapd to a sirup, to which the above test is applied. To est glycerol present as glycerides, glycerophosphate or lecithin, the material must first be hydrolyzed with NaOH. The coloration given with thiosulfate appears to be due to the transitory formation of the unstable blue modification of S. The test may be used also for the detection of thiosulfate.

II C A

New analytical applications of the reaction of ammonia on resorcinol in the presence of cations. LIGOR REY *Bull soc chim* [4], 47, 1192-3(1930); cf *C A* 23, 1989, 24, 2397.— NH_3 was added to a 5% soln of resorcinol in ether, just to the point of resolution. A soln contg 0.0011% $CdCl_2$, $2H_2O$, added carefully, gave a blue ring at the junction of the 2 layers in 26 min. It persisted 5-10 min. Lower concns, gave a noncharacteristic violet. In the analysis for Sn, stannic Sn gave the best results. The Sn was oxidized with Br until the soln was a light yellow. Two or 3 cc, was shaken with 5 cc, NH_3 (d = 0.925) for several min., 2 or 3 cc of 5% aq resorcinol was added, giving a white ppt of $Sn(OH)_2$, which turns blue in 2-20 min. This test is sensitive to 0.0015% $SnCl_2$, $2H_2O$. Lower concns gave a noncharacteristic gray. The result is not affected by less than 400 times as much Sb as Sn.

G M P

Detection and estimation of hydroxyl groups in organic compounds. HUFFMANN, *Pharm. Ztg* 76, 113-4(1931).—This is a commentary on the usual methods (by the aid of $AcCl$, Ac_2O and $BzCl$) in connection with expts on resorcinol, nreitol, Me salicylate and 2,7-hydroxytetralol, in which the procedure of Verky and Bölling was used.

W. O. E.

Iodoxybenzoate as a test reagent for free phenolic hydroxyl groups in organic compounds. CHANCEV D. LEAKE *Proc. Soc. Exptl Biol. Med* 28, 148-50(1930).—Morphine salts treated with an aq soln of NH_4 iodoxybenzoate give a straw to garnet color reaction. No other opium alkaloid, except apomorphine, gives a color with this reagent. The reagent does not decompose the compds with which it comes in contact; it reveals by an oxidizing color reaction the presence of a "free" phenolic hydroxyl group. The test should be useful in detecting and differentiating morphine and apomorphine from other opium alkaloids, adrenaline from ephedrine, acetarone from other quinquevalent arsenicals. It detects phenol and cresol as preservatives in org fluids. The test substance need not be miscible with the reagent; thorough shaking is sufficient. The pharmacol effect of substances treated with this reagent does not

group to be altered. Substances such as codeine, in which the H of the phenolic hydroxyl group is replaced, have much less toxicity and general pharmacol action than the parent substances. C. B. BAILEY

Detection of hydrogen in volatile organic compounds. LAWRENCE STONE. *Chemist Analyst* 20, No. 2, 7(1931).—By passing the vapors through a small opening so that they strike against hot CuO in a test tube, water is formed and condenses on the upper walls of the tube. W. T. H.

Note on the Ter Meulen-Heslinga methods for the estimation of nitrogen, carbon and hydrogen in organic material. E. P. GAFFING AND C. L. ALSBRAG. *J. Am. Chem. Soc.* 53, 1037-4(1931).—This is a discussion and suggestion of modifications of the methods for N, C and H developed by Meulen and Heslinga (cf. *C. A.* 19, 226, 2465). C. J. WEST

A method for the (mercurimetric) determination of formaldehyde and hexamethylenetetramine. ENRIQUE E. REBAGLIATI. *Ann. Farm. Bioquim.* 1, 150-8(1930).—The substances are reduced with the aid of Knapp's reagent modified, and the final detn is accomplished by titration with 0.1 N AgNO_3 according to the method of Denigès. R. claims that the method is rapid and exact. B. S. LEVINE

New method for the estimation of halogens in organic substances. JUAN A. SANCHEZ. *Ann. farm. bioquim.* 1, 121-34(1930).—Heat the org. substance in a closed tube together with solid K_2MnO_4 ext. with water and destroy the excess of the reagent with H_2O_2 . Detn. the halogen titrimetrically. The method can be applied to quantities of org. matter as low as 0.02 g. It is applicable with a high degree of precision to aliphatic and cyclic derivs. and to substances rich in C and poor in halogen, and vice versa. The detn. of halogens in substances such as C_2Cl_4 and the halogen derivs. of phenyl is accomplished with greater ease than by the method of Carus. B. S. LEVINE

The activity of Professor F. Wald in analytical chemistry (QUADRAT) 2. Conductivity measurements and titrations (HANN) 2. Oxidation of heterocyclic As compounds by I (RAZUVAYEV, MALINOVSKII) 10.

Fälschungen. Edited by SIEGFRIED TCKEWL. Graz: Ullr Moser. 93 pp. M. 12.
Reviewed in *Analyst* 56, 141(1931).
SCHUFFTAN, P. Gasanalyse in der Technik. Leipzig: S. Hirzel. 80 pp. M. 5.50

8—MINERALOGICAL AND GEOLOGICAL CHEMISTRY

EDGAR T. WHERRY AND J. F. SCHAIKEN

Gallium. IV. Occurrences of gallium in zinc minerals. JACOB PAPISH AND CHESTER B. STILSON. *Am. Mineral.* 15, 521-7(1930).—The arc spectrographic method, with graphite and also Ag electrodes, was used in a study of a no. of Zn minerals for the detection of Ga. The relative intensity of the Ga lines is listed. Sphalerite, gahnite, hopeite, parahopeite and adamite were found to be galliferous. Ga was not detected in calamine, smithsonite, zincite, hydrozincite, goslarite, hodgkinsonite, hardystonite and clinohedrite. A. M. BRANT

Tellurium and tellurium minerals of Hungary. S. KOCH. *Bányász. Kohás. Lapok* 62, 425-9, 449-52(1929).—Te is the one element discovered by a Hungarian scientist. Pál Kitaibel. Eleven Te minerals are found in Hungary, ten of which were first discovered there. The genetic relations of these are summarized. Native Te of Fácshánya contained 0.15-2.78% Au. Tetradymite, hessite, wehrhite, petzite, muthmannite, krennerite, sylvanite, nagyágite, tellurite and stützite are described. The mines which once yielded such Au-contg. tellurides belong now to Roumania, and most of them are not being worked. S. S. DE FINAII

Notes on rickardite, a new occurrence. WM. P. CRAWFORD. *Am. Mineral.* 15, 272-3(1930).—Rickardite, Cu_2Te , has been found in the sulfide ores of the Briggs' mine, Warren, Ariz. A. M. BRANT

Crystallography of stibnite and orpiment from Manhattan, Nevada. CHARLES PALACHE AND DAVID MODELL. *Am. Mineral.* 15, 365-74(1930).—The crystallographic features are given in detail. The data on orpiment give addnl. proof that the mineral is monoclinic. A. M. BRANT

Crystal habit of zinc blende. G. KALB AND L. KOCH. *Centr. Mineral. Geol.* 1929A, 353-7. —Zn blende is referred to the octahedral isoharmonic crystal type, 2 forms, resp., black and reddish brown to reddish yellow, are distinguished. B. C. A.

Arsenoferrite from Jachymov, Czechoslovakia. WM. F. FOSHAG AND M. N. SHORT. *Am. Mineral.* 15, 423-9(1930).—This mineral, FeAs_2 , is the arsenide member of the pyrite group. It is isotropic in character, differing from löllingite which is anisotropic. It can be distinguished from smaltite only by chem. tests. The color is dull medium gray. The phys. properties are: metallic luster, no cleavage, hardness 5.5, sp. gr. 6.42. Jachymov arsenoferrite occurs as fine-grained irregular masses in a carbonate gang. Chem. analysis of material contg. some gang gave: Fe 24.88, Cu 1.34, Pb 0.05, CaCO_3 34.00, MgCO_3 1.57, As 66.84, S 1.03, sum 99.78%. Co and Ni were not detected. The blowpipe characteristics are the same as those for löllingite. A. M. BRANT.

The crystal structure of cryolithionite. G. MEYER. *Z. Krist.* 75, 265-87(1930).—A detailed report of the structure of $\text{Li}_2\text{Na}_2\text{Al}_2\text{F}_{10}$ (cf. C. A. 22, 2340). The side of the unit cube is 12.097 ± 0.004 Å and the calcd. d is 2.791. Coordinates for the F atoms are 0.035, 0.04 and 0.64 (all ± 0.01). Although the structure is analogous to that of garnet, the diffraction effects are quite different, for in garnet the O atoms are almost negligible, while in cryolithionite the corresponding F atoms are the most important. L. S. RAMSDELL.

Boleite, pseudoboleite and cumengeite. B. GOSSNER. *Z. Krist.* 75, 365-7(1930).—A criticism of the structure detn. of Friedel and Hocart (C. A. 24, 4734). G. maintains that the x-ray data show boleite to be cubic and not tetragonal. L. S. RAMSDELL.

Origin, weathering and artificial coloring of agate. H. HEINZ. *Chem. Erde* 4, 501-25(1930).—The different layers in agates were examd. chemically, optically and by x-rays. Analyses are given of the white weathered crust and of the fresh agates; e. g., flint from the English chalk gave I for the inner black portion and II for the white chalk-like crust.

	SiO_2	$(\text{Fe,Al})_2\text{O}_3$	$(\text{Mg,Mn})\text{O}$	CaO	Ign.	D
I	97.93	0.27	0.20	0.59	1.25	2.506
II	96.72	0.28	0.09	1.71	1.68	2.501

The portion sol. in 10% KOH soln., taken to represent admired opal, varies in different bands from 5.02 to 30.69%, with a corresponding range in d from 2.620 to 2.601. Artificial coloring, by soaking in sugar soln. for 14 days and then in concd. H_2SO_4 , was effected in the finer-grained layers contg. more opal, rather than in porous weathered layers. Banded and tubular structures, similar to those seen in agates, were obtained experimentally by the interaction of silicic acid sols and salt sols. B. C. A.

Agate. G. LRYCK AND H. HEINZ. *Chem. Erde* 4, 526-8(1930).—A discussion of the preceding abstr. The banding of agate is attributed to the alternating action of sols of silicic acid and of salts, the former during the dry season and the latter during the wet season of the year. B. C. A.

The relations of magnetite and ilmenite in the magnetite deposits of the Duluth gabbro. G. M. SCHWARTZ. *Am. Mineral.* 15, 243-52(1930).—Polished surfaces of the titaniferous magnetite ores and gabbro were examd. in reflected light. The Ti was found to occur only in the form of ilmenite, as grains and as intergrowths along the parting planes of magnetite. The intergrowths are estd. to amount to 15% of the ore. Their texture is attributed to the breaking down of a solid soln. A. M. BRANT.

Four stages in the alteration of the Villeneuve uraninite. H. V. ELLSWORTH. *Am. Mineral.* 15, 455-60(1930).—The specimen, found in 1886, shows well-marked zones of material due to progressive alteration. The central zone A has a steely blue-black color, hardness 6, sp. gr. 9.144, zone B has a pitch-black color, hardness 4.5, sp. gr. 7.779; zone C has a bright "flame-scarlet" color, hardness 3.5, sp. gr. 5.273, waxy luster; zone D is a very thin irregular layer with hardness 3 and has a buff to apricot-yellow color and waxy luster. The chem. analysis of the first 3 are tabulated. The results indicate that when analyzing altered uraninites for age detns., all red and yellow alteration products should be excluded. This is due to the abnormal increase in the Pb-Ur ratio in these products. A. M. BRANT.

Uraninite from Placer de Guadalupe, Chihuahua. ROGER C. WELLS. *Am. Mineral.* 15, 470-3(1930).—The mineral occurs in a calcite vein about 3 ft. thick between porphyry and shale. It is intimately associated with Au. The nodules range from small specks to the size of a walnut. Above the water level the pitchblende is leached out leaving wires of Au and small quantities of uranophane and a dark green U mineral. The mean of 2 analyses is: UO_2 70.09, UO_3 22.69, ThO_2 0.20, CeO_2 0.71, La_2O_3 etc., 1.02, Y_2O_3 etc., 3.41, FeO 0.10, As_2O_3 0.06, PbO 0.40, CaO 0.30, TiO_2 0.06, Al_2O_3 0.23, H_2O (0.41), sum 99.70%. The age of the veins calcd. from the Pb-Ur ratio

is 36 million years. The sp. gr., 10.63, detd. by pycnometer, is the highest so far reported for uraninite. A. M. BRANT

Eubedral magnesite crystals from Winkler County, Texas. JOHN T. LONSDALE. *Am. Mineral.* 15, 238-9(1930).—Magnesite crystals 1-5 mm long were obtained from drill cores at 2690 feet. The matrix is dolomite limestone contg. shint nodules. Many of the crystals are stained with petroleum or asphalt. A. M. BRANT

The occurrence of magnesite and of iron ores at Tarrekalsse, Sweden. ALVAA HÖGBOOM. *Geol. Fören. Förh.* 52, 417-30(1930).—The magnesite is usually accompanied by chlorite, asbestos and magnetite. The probable formation of the ores is discussed. WILHELM SEGERBLÖM

Feldspars in the Adirondack anorthosite. HAROLD L. ALLING. *Am. Mineral.* 15, 267-71(1930).—A discussion of an article by Barth (*C. A.* 24, 6258). A. concludes that (1) the plagioclase of the Adirondack anorthosites consists of labradorite grading into andesine with an appreciable though small percentage of K feldspar in solid soln., and (2) late-stage interstitial feldspars are Na-rich anorthoclases and K-bearing oligoclases. The latter may be untwinned. A. M. BRANT

A chromium-bearing pyroxene from Jagersfontein, S. Africa. J. O'DANIEL. *Z. Krist.* 75, 578(1930).—This dark green mineral occurred as a nodule in the blue-ground. The angle between the 110 and $\bar{1}\bar{1}0$ cleavages was $92^{\circ}49'$; between 110 and 001, $79^{\circ}7'$. Analysis gave SiO_2 53.53, Al_2O_3 1.30, Cr_2O_3 1.96, FeO 2.10, CaO 22.96 and MgO 17.88, sum 99.73%, $n_D^{20} = 1.6722$, $n_F^{20} = 1.6817$, $\gamma = 1.7015$. L. S. RAMSDALL

A titaniferous augite from Chaudrawati, Sirohi State, Rajputana. A. L. COULSON. *Records Geol. Survey India* 63, 448-50(1930).—The mineral occurs in a contact metamorphic product at the junction of olivine gabbro with calcic rock. Its associated minerals are calcite, wollastonite, pectolite, quartz, microcline, orthoclase, plagioclase and magnetite. Optical data are given. ALFRED H. TWEED

Iron-rhodonite from Tuna Håstberg. N. SUNDSTRÖM. *Geol. Fören. Förh.* 52, 401-6(1930).—An ore found at Tuna Håstberg was a very Mg-poor olivine holding 27.0% MnO and 40% FeO, small grains of garnet and a pyroxene-like mineral which made up 4 to 5% by vol. of the ore. The last named mineral was found to be an Fe-rich rhodonite with the compn. SiO_2 47.78, Al_2O_3 0.08, Fe_2O_3 0.11, FeO 14.51, MnO 29.20, MgO 1.93, CaO 6.55, H_2O 0.09, sum 100.25%. The mineral is colorless or slightly pink and has 2 excellent cleavages nearly at right angles to each other. Discussion of the work of Weibull and of Palmgren and Sobral on similar minerals leads S. to conclude that the name iron rhodonite should be reserved for members of the series which contain considerable amts. of FeO. WILHELM SEGERBLÖM

The structure of anthophyllite, $\text{H}_2\text{Mg}_3(\text{SiO}_3)_8$. D. E. WARREN and D. I. MOWELL. *Z. Krist.* 75, 161-78(1930)(in English).—Anthophyllite is the simplest orthorhombic amphibole. X-ray data lead to the formula $\text{H}_2\text{Mg}_3(\text{SiO}_3)_8$ rather than $\text{Mg}_3\text{Si}_3\text{O}_{10}$. Rotation photographs give $a = 18.6$, $b = 17.9$ and $c = 5.27$ Å. U. The unit contains 4 mols. and the space group is V_2^{16} . Compared with enstatite (*C. A.* 25, 1420) the values for a and c are identical, but b for anthophyllite is twice that for enstatite. A block of enstatite is multiplied by the action of symmetry planes to form the anthophyllite cell, except for $1/8$ of the Mg atoms, which fall in unlikely positions and drop out. Each Si is surrounded by 4 O atoms tetrahedrally arranged. Sharing of tetrahedra corners gives rise to the characteristic amphibole double Si-O chains. One twelfth of the (O, OH) positions are not bound to Si, but to only 3Mg. These are only singly charged and must be OH. The general relations between amphiboles and pyroxenes are discussed. L. S. RAMSDALL

Soda-rich anthophyllite asbestos from Trinity County, California. J. D. LAUDERMILK and A. O. WOODFORD. *Am. Mineral.* 15, 259-62(1930).—The country rock is serpentine contg. minute branching veins of chrysotile and much thicker veins (up to 5 cm) of amphibole asbestos. They are short branching and lie in the crushed slickensided rock in roughly parallel positions. The fibers are usually perpendicular to the walls. The material was found to be anthophyllite with $a = 1.606$, $b = 1.613$, $\gamma = 1.623$. Z, parallel to elongation, was pale yellow green, in the other directions colorless. Analysis gave SiO_2 87.79, Al_2O_3 2.03, Fe_2O_3 trace, FeO 5.32, MgO 21.12, CaO 5.10, Na_2O 7.40, K_2O absent, MnO trace, $\text{H}_2\text{O}^+ 1.80$, $\text{H}_2\text{O}^- 0.30$, sum 100.74%. A. M. B.

An association of anthophyllite and enstatite. F. COLES PHILLIPS. *Geol. Mag.* 67, 513-6(1930).—These minerals were observed in thin sections of rocks from the Shetland Islands. The anthophyllite shows a variable degree of alteration to talc. The enstatite is of later formation than the anthophyllite, but there is little evidence that the pyroxene formed at the expense of the amphibole. A. M. BRANT

Hypersthenization and its explanation. D. GUIMARAES. *Ann acad. brasil sci.* 2, 1-11(1930) —The formation of hypersthene from diopside by way of clinoenstatite is discussed from the crystallographic point of view, and a general theory is put forward to account for the chem transformations which have occurred for silicate minerals. It is supposed that at high temps and pressures the small quantity of water present acts as an electrolyte, and interaction of silicates, or of silicates and SiO_2 , and the transformation of one mineral into another is possible in directions controlled by the H-ion concn. B C A

Tremolite from near Jasidih, Bihar. A L COULSON *Records Geol Survey India* 63, 444-6(1930) ALDEN H EMERY

Changes in hornblende at about 800°. VIRGIL E BARNES. *Am. Mineral.* 15, 393-417(1930) —A set of 14 samples of hornblende including edenite and pargasite and 6 of the tremolite-actinolite series were studied optically before and after heating in air. The samples were later heated to 850° in a stream of H and the optical properties again studied. The results are given in tabular and graph form. It was found that green hornblende when heated in air changes to brown hornblende, with the optical properties of basaltic hornblende. The Fe poor hornblendes do not change on heating. Tremolite behaves like the Fe-poor hornblendes and actinolite like the iron-rich ones. The changes observed in heating the latter are: an increase in d_{ns} , birefringence and dispersion; a decrease in extinction angle, a change in optic angle to about 65°; a change in pleochroism from greens to browns. The birefringence of brown hornblende in violet is greater than that for red, and the opposite is in general true for common hornblende. The wt.-% loss on heating is only a fraction of the H_2O obtained by analysis; there is marked conversion of Fe^{++} to Fe^{+++} . Heated hornblendes in general show a more complete change in Fe^{++} than basaltic hornblendes. Dehydration has little effect on the optical properties, but they change almost in proportion to the change in the total Fe oxides, with Fe^{+++} being especially influential. Brown hornblende when heated in H changes back to green hornblende. Inversion does not take place. In dehydration, H and not H_2O (except that which is not a constituent part of the space lattice) is given off. The O remains in the mineral, either oxidizing Fe^{++} to Fe^{+++} , or in the absence of Fe^{++} remaining free because of its size. A M. BRANT

Notes on some Franklin minerals. L. H. BAUER AND HARRY BERMAN. *Am. Mineral.* 15, 340-8(1930) —*Cummingtonite* (Zn, Mn-bearing) was found embedded in rhodonite or garnet. The green prismatic crystals have a cleavage angle of $54^\circ 28'$. The chem compn is: SiO_2 49.74, Al_2O_3 1.72, FeO 12.80, MgO 8.31, ZnO 10.46, MnO 18.79, CaO 0.49, Na_2O 0.22, H_2O 2.16; sum 99.69%. The optical properties are: biaxial-, $2V = 75^\circ$, $ZAc = 15^\circ$, $Y = b$, $\alpha = 1.657$, $\beta = 1.674$, $\gamma = 1.685$. *Apophyllite* found as pale-pink radiating crystals in limestone has the compn SiO_2 50.90, CaO 24.74, K_2O 3.70, Na_2O 0.42, MnO 0.47, ZnO 1.79, H_2O 17.71%. The derived formula is $(\text{Ca}, \text{K}, \text{Na}, \text{Mn}, \text{Zn})_2 \text{Si}_4\text{O}_{13} \cdot 3\frac{1}{2}\text{H}_2\text{O}$. *Barynite* occurs as thin films and veinlets assocd with garnet, axinite, hardystonite. The chem compn is: SiO_2 16.84, PbO 77.35, MnO 3.33, FeO 0.23, Al_2O_3 0.59, CaO 0.21, MgO 0.78, ZnO 0.30, H_2O 0.07, sum 99.70%, the formula is $3\text{PbO} \cdot 2\text{SiO}_2$. The optical properties are: uniaxial-, $\omega = 2.033$, $\epsilon = 2.015$. *Celesite* is reported from Sterling Hill. It occurs in cavities in franklinite ore as good crystals tabular parallel to $a(100)$. *Chlorophoenicite* (previously described as clinozonite) is compared with that mineral. *Dalolite* has been found in crystals tabular to $b(010)$. *Ferroschallerite*, in which the Mn of schallerite is partially replaced by Fe, has the compn: SiO_2 31.12, MnO 29.22, FeO 17.12, MgO 0.12, ZnO 3.63, As_2O_3 12.46, H_2O 6.42, sum 100.09%. A single specimen of *manganbrucite* was found. This had a Mg to Mn ratio of 5:1. The chem analysis of *svabite* is: SiO_2 0.29, H_2O 1.32, F 1.41, CO_2 trace, CaO 45.89, ZnO 1.54, MnO 1.23, MgO 0.84, PbO 0.51, As_2O_3 35.24, P_2O_5 12.54, sum, corrected for O = F, 100.22%; sp gr is 3.542 and $\omega = 1.684$. The relations of the members of the apatite-svabite series are discussed. A. M. BRANT

The structure of hardystonite, $\text{Ca}_2\text{ZnSi}_2\text{O}_7$. B E WARREN AND O. R. TRAUTZ. *Z. Krist.* 75, 525-8(1930)(in English) —From rotation and oscillation photographs it is shown that the unit cell and space group of hardystonite are the same as those found for melilite (C. A. 24, 4736). The dimensions are $a = 7.83$, $c = 4.99$ A U, $c/a = 0.637$. The at. arrangement is the same as in melilite, except that the (Mg, Al) is replaced by Zn. L. S RAMSDELL

Crystal lattice of eulytite, $\text{Bi}_2\text{Si}_2\text{O}_7$. G. MENZER. *Centr. Mineral. Geol.* 1928A, 420-1. —Powder photographs show a cubic, space-centered lattice, $a_0 = 10.25 \pm 0.02$ A. U.; $d_{632} = 0.21$. If a mol. wt. of 1109 is assumed, there are 3.7 mols in the unit cell. B. C. A.

The structure of danburite, $\text{CaB}_2\text{Si}_2\text{O}_{10}$. C. DUNBAR AND F. MACHATSCHKE. *Z. Krist.* 76, 133-46(1930)(in English).—The orthorhombic unit cell has the dimensions $a = 8.75$, $b = 8.01$ and $c = 7.72$ A. U. If the symmetry is holohedral, the space group is V_{16}^{11} . There are 4 mols. in the cell. The unit of structure is not an SiO_4 group, but rather Si_2O_7 , 2 tetrahedral groups being linked by a common O atom. The eighth O atom forms a tetrahedron with O atoms from $3\text{Si}_2\text{O}_7$ groups. The B atoms are at the centers of these tetrahedra. Although the presence of B makes the detn. of structure difficult, parameters are derived for all positions. L. S. RAMSDELL.

Zoisite and its native rock from the high mountains of Västernorrland. ERLAND GARR. *Bull. Geol. Inst. Univ. Uppsala* 22, 289-98(1930).—Three samples of amphibolite from different localities were examd. The rock is dark grayish green, somewhat slate-like, with porphyritic injections of zoisite in a granular base of hornblende, plagioclase and quartz. The long, flat, striated zoisite needles are white with a tinge of green, and are, in general, oriented with respect to the cleavage plane (100). The optical orientation is $\alpha = c$, $\beta = b$, $\gamma = a$. The sp. gr. is 3.376, it is optically + and $2V = 50-56^\circ$, $n = 1.7011$, $\beta = 1.7020$, $\gamma = 1.7073$. Analysis gave SiO_2 39.37, Al_2O_3 32.06, FeO 1.47, FeO 0.15, MnO 0.02, CaO 24.32, MgO none, TiO_2 0.56 and H_2O 1.89%. Descriptions and analyses of the other mineral constituents are included. The composite material from Spångsjället, sp. gr. 3.090, contained hornblende 70.7, zoisite 10.3, quartz 7.7, plagioclase 3.5, garnet 5.7 and titanite 2.1% by wt. W. L. HILL.

Sapphirine in the Vinsgapatam district. H. CROOKSHANK. *Records Geol. Survey India* 63, 446-5(1930).—The material occurs as a thin rim around grains of hercynite. Optical data are given. ALDEN H. ECKERT.

Formulation of silicates. W. KRYTZE. *Chem. Erde* 4, 549-50(1930).—A further discussion of the chem. formula of tourmaline, with comparison of the views of E. (C. A. 25, 1461) and of Machatschke (C. A. 23, 5132 and C. A. 25, 1461). To conform with the detn. that the unit cell contains 3 mols., the empirical formula may be multiplied by $3/2$ or $9/4$. B. C. A.

The space lattice of natrolite. FRANK HALLA AND ERNST NEHL. *Z. Krist.* 75, 421-9(1930).—The orthorhombic unit cell of natrolite has the dimensions $a = 18.284 = 0.004$, $b = 18.715 = 0.027$ and $c = 6.632 = 0.008$ A. U. It contains 8 mols. of $\text{NaAl}_2\text{Si}_2\text{O}_7 \cdot 2\text{H}_2\text{O}$. If holohedral symmetry is assumed, the space group is V_{16}^{11} ; otherwise it is C_{2v}^{15} or V_7 . L. S. RAMSDELL.

The crystal structure of muscovite. W. W. JACKSON AND J. WEST. *Z. Krist.* 76, 211-27(1930)(in English).—The structure for muscovite proposed by Pauling (C. A. 24, 2694) has been confirmed, and the details of the structure are discussed. Four mols. of $\text{KAl}_2(\text{AlSi}_3\text{O}_{10})(\text{OH})_2$ are in the monoclinic unit, with $a = 19.96$, $b = 9.02$, $c = 5.18$ A. U. and $\beta = 95^\circ 30'$. The close relations between diopside, tremolite and muscovite are discussed. L. S. RAMSDELL.

The structure of chrysotile, $\text{H}_3\text{Mg}_3\text{Si}_2\text{O}_{10}$. B. E. WARREN AND W. L. BRAGG. *Z. Krist.* 76, 201-10(1930)(in English).—Chrysotile (variety of serpentine) occurs only in bundles of parallel fibers, with random orientation about the fiber axis (= c axis). It is not possible to obtain rotation photographs about any other axes, and exptl. data are consequently meager. This proposed structure is believed, however, to be correct in its main features. The unit cell is monoclinic, with $a = 14.66$, $b = 18.5$, $c = 5.53$ A. U. and $\beta = 93^\circ 16'$. The space group is probably C_{2v}^{15} . The unit contains 4 mols. of $(\text{OH})_3\text{Mg}_3\text{Si}_2\text{O}_{10} \cdot \text{H}_2\text{O}$. Each Si atom is surrounded by a tetrahedron of O atoms. Sharing of O atoms by adjacent tetrahedra gives rise to double Si-O chains, just as in the amphiboles. These have the compn. Si_2O_7 , and between them lie sections of Mg and OH arranged as in brucite ($\text{Mg}(\text{OH})_2$). Although the formula calls for $4\text{H}_2\text{O}$, these are not H_2O mols., as each H is shared by 2 O atoms. The lateral binding between the chains is very weak, the perfect fibrous character of chrysotile being thus accounted for. L. S. RAMSDELL.

Aluminum silicate from Althar, S. Serbia. F. TRCAN. *Bull. soc. franc. minéral.* 52, 42-7(1929).—The product, having a compn. corresponding with the formula $\text{Al}_2\text{O}_3 \cdot 3\text{SiO}_2 \cdot 7\text{H}_2\text{O}$, has dm 1.9818, when heated it evolves water (completely at 650°) having an acid reaction. The dehydration curve (CaCl_2 , H_2SO_4 , or P_2O_5 being used) is continuous. B. C. A.

The chemical composition of lessingite and cerite from the Kyshtym district. V. N. KUPROVICH. *Trans. Inst. Econ. Mineral. and Met. (Moscow)* No. 44, 46(1930).—The formula for the new mineral lessingite may closely approach $\text{H}_2\text{Ca}_2\text{Ce}_2\text{Si}_2\text{O}_{10}$. On the assumption of the presence of an admixt. of bastnäsite, the compn. of the cerite is

$4\text{Ce}_2\text{O}_3 \cdot 6\text{SiO}_2 \cdot \text{H}_2\text{O}$ which closely approaches Swedish cerite but contains less H_2O

H. C. PARISH

Cenosite from North Burgess Township, Lanark County, Ontario. R. P. D. GRAHAM AND H. V. ELLSWORTH *Am Mineral* 15, 205-19(1930).—The mineral occurs as minute pink or pale rose colored crystals on calcite and quartz in soln. cavities in a coarse granular limestone. Chalcopyrite and celestite are also present. It is orthorhombic in crystal, prismatic in habit. The mineral is optically —, with 2V about $39^\circ 45'$, $\alpha = 1.644$, $\beta = 1.689$, $\gamma = 1.691$. Orientation $a = \alpha$, $b = \beta$, $c = \gamma$. Other phys. properties are hardness 5-6, sp gr 3.612, vitreous to pearly luster, no cleavage. Chem. analysis gave SiO_2 34.66, $(\text{Ce}, \text{La}, \text{Dy})_2\text{O}_3$ 3.22, $(\text{Yt}, \text{Er})_2\text{O}_3$ 35.46, CaO 16.72, SrO 0.31, Al_2O_3 , Fe_2O_3 , BeO 0.22, MnO 0.02, Na_2O 0.27, K_2O trace, MgO 0.19, CO_2 6.58, SO_3 0.04, H_2O 2.54, H_2O (110-130°) 0.04, sum 100.23%, formula deduced, $2\text{CaO} \cdot (\text{Ce}, \text{Yt})_2\text{O}_3 \cdot \text{CO}_2 \cdot 4\text{SiO}_2 \cdot \text{H}_2\text{O}$

A. M. BRANT

Allodelphite, a new silicoarsenite from Långban. P. QUENSEL AND H. VON ECKER-MANN *Geol. Foren. Forh.* 52, (139-46)(1930).—The mineral which Flinck labelled "un described mineral no. 325" in his catalog of new minerals from Långban (cf. *C. A.* 19, 2623, 20, 1778) is now named allodelphite to show its relation to synadelphite. Allodelphite occurs as thin crusts of small well defined elongated tabular crystals of a dark red brown color, but deep wine red in transmitted light. The luster is submetallic, fracture conchoidal and streak chocolate brown. It is optically —, sp gr = 3.573. The compn. is SiO_2 6.23, As_2O_3 21.91, As_2O_5 0.62, Sb_2O_3 0.15, Al_2O_3 1.50, Fe_2O_3 0.98, PbO 0.39, MnO 50.30, CaO 1.10, MgO 6.22, K_2O 0.74, Na_2O 0.53, H_2O 8.82, sum 99.49%. Allodelphite is evidently a new representative of the small group of silicoarsenites or silicoarsenates. Two general formulas are suggested $9\text{RO} \cdot \text{As}_2\text{O}_3 \cdot \text{SiO}_2 \cdot 5\text{H}_2\text{O}$ and $5\text{RO} \cdot 2\text{R}_2\text{O}_3 \cdot \text{As}_2\text{O}_3 \cdot \text{SiO}_2 \cdot 5\text{H}_2\text{O}$. Q and E discuss these in the light of a new analysis of the related mineral synadelphite and conclude that, pending further investigation, the formula of allodelphite may be assumed to be $\text{MnO} (\text{Mn}_2, \text{As}, \text{MnSi})_2\text{O}_3 \cdot \text{H}_2\text{O}$. W. S.

The hiddenite deposit in Alexander County, North Carolina. CHARLES PALACHE, S. C. DAVIDSON AND E. A. GORANSON *Am Mineral* 15, 280-302(1930).—The genetic history of this deposit was traced, 3 stages of mineralization being distinguished, as follows: (1) the *lit-par-lit* pegmatite stage, (2) the hiddenite pegmatite stage, (3) the hiddenite cavity stage. The minerals of the latter stage are described in some detail. The optical properties of *holmquistite* are $\alpha = 1.625$, $\beta = 1.645$, $\gamma = 1.654$; pleochroism X = light yellow, Y = violet, Z = purplish violet; Z nearly parallel to c , dispersion weak $\rho > \nu$, $2V = 51^\circ$. The mineral is compared as to chem. compn. with the find at Uto, Sweden.

A. M. BRANT

Paragenesis of the Newry pegmatite, Maine. H. J. FRASER. *Am Mineral* 15, 349-64(1930).—Structural and textural data indicate 3 periods during the formation of the pegmatite, and that the process was essentially continuous. In the first period mainly silicates were deposited. During the second period these changed to Li silicates and phosphates. The third period was one of deposition of Mn, Fe, Al and Be phosphates.

A. M. BRANT

Pegmatite minerals of Poland, Maine. HARRY BERMAN AND F. A. GONYER. *Am Mineral* 15, 375-87(1930).—The minerals found in pockets in microcline comprised *amblygonite*, as a compact rim around the pockets, *lithophilite*, partially replaced by *reddingite*. The chem. compn. is FeO 10.96, MnO 31.90, Na_2O 0.30, Li_2O 9.55, H_2O + 0.40, P_2O_5 46.35, insol 0.16, sum 99.63%. The optical properties of the analyzed material were biaxial+, $2V = 60^\circ$, α and γ parallel to cleavages $\alpha = 1.675$, $\beta = 1.679$, $\gamma = 1.688$. *Rhodochrosite* occurs as an irregular layer over the *amblygonite* containing cavities with small protruding crystals of the mineral, strongly etched. *Quartz* occurs as min. doubly terminated crystals in veinlets in the *rhodochrosite*. *Eosphorite* (cf. *Drumman C. A.* 10, 31) occurs as well formed crystals embedded in *rhodochrosite*. *Reddingite* replaces *rhodochrosite* as granular masses and as veinlets in *lithophilite*. The color varies from colorless to deep brownish red, depending on the degree of alteration to *landesite*. The chem. compn. is Fe_2O_3 0.95, FeO 12.68, MnO 38.36, CaO 0.15, Na_2O K_2O trace, H_2O + 13.16, P_2O_5 34.52, insol 0.45, sum 100.27%, formula $3(\text{MnFe})\text{O} \cdot \text{P}_2\text{O}_5 \cdot 3\text{H}_2\text{O}$, with $\text{Mn}:\text{Fe} = 3:1$, optically biaxial+ with $2V = 65^\circ$; $\alpha = 1.655$, $\beta = 1.662$, $\gamma = 1.683$. *Dickinsonite* occurs as small, tabular, highly etched crystals associated with *fairfieldite*. The color is brownish to yellowish green. The chem. compn. is FeO 12.33, MnO 31.83, MgO 1.67, CuO 2.01, Na_2O 7.41, K_2O 1.73, Li_2O 0.20, H_2O 1.82, P_2O_5 40.78, insol 1.00, sum 100.78%; formula $7(\text{MnFe})\text{O} \cdot 2(\text{Na}_2\text{K}_2\text{Ca})\text{O} \cdot 3\text{P}_2\text{O}_5 \cdot \text{H}_2\text{O}$, optically biaxial+ with $2V$ near 90° ; $\alpha = 1.648$, $\beta = 1.655$, $\gamma = 1.662$. *Fairfieldite* occurs as white foliated plates in *rhodochrosite* and on *dickinsonite*. Its compn. is FeO 4.74, MnO 14.82, CaO 30.85, Na_2O 0.41, K_2O none, H_2O 9.70, P_2O_5 39.55, insol

0.50, sum 100.58%; formula $(\text{MnFe})\text{O}_2\cdot 2\text{CaO}\cdot \text{P}_2\text{O}_5\cdot 2\text{H}_2\text{O}$, with $\text{Mn}:\text{Fe}::3:1$; optically biaxial + with $2V = 86^\circ$; $\alpha = 1.640$, $\beta = 1.650$, $\gamma = 1.660$. *Apatite* occurs as small "nailhead" crystals resting on the other crystals. *Laudenite*, a new mineral, is a brown alteration product of reddingite. The crystals are rough and of a form suggesting the unit pyramid of the latter. Two cleavages are shown. The chem. compn is $\text{Fe}_2\text{O}_3 + \text{FeO } 13.91$, Mn_2O_3 2.69, $\text{MnO } 33.65$, $\text{MgO } 3.07$, $\text{CaO } 1.39$, $\text{H}_2\text{O } 13.60$, P_2O_5 31.94, insol 0.13, sum 100.38%. This analysis yields the formula $3\text{Fe}_2\text{O}_3\cdot 20\text{MnO}\cdot 8\text{P}_2\text{O}_5\cdot 27\text{H}_2\text{O}$. The mineral is related to salmonite in compn and to sickleite in optical properties. These are biaxial -, $2V$ large; Z perpendicular to best, and X to an inferior cleavage; pleochroism $X = \text{dark brown}$, $Y = \text{light brown}$, $Z = \text{yellow}$, $\alpha = 1.720$, $\beta = 1.723$, $\gamma = 1.735$. *Cassiterite* is contemporaneous with rhodochroite. It is identified optically by the following: uniaxial +; $\omega = 2.01 = \text{light green}$, $\epsilon = \text{brown}$. The pleochroism is worthy of note. A *beryl* crystal from this locality has on the base an excellent etch figure of the forms $(10\bar{1}1)$ and $(11\bar{2}1)$. Its $\omega = 1.585$, indicating high alkali content and probably Cs . This crystal was not a part of the phosphate suite of minerals and is probably of an earlier period. The solns leading to the formation of the foregoing minerals were essentially of Mn phosphate with carbonate and water, Li rich in the earlier phase and Ca rich toward the end. A. M. BRANT

Pegmatite minerals of Ontario and Quebec. HUGH S. SPENCE *Am Mineral.* 15, 430-50, 474-96 (1930).—The minerals occurring in (1) the normal granite pegmatites of Ontario and Quebec and (2) the syenite pegmatites of the Walberforce district, Ontario, are reviewed and the pegmatites described. The minerals are listed for those of more restricted regional development: (1) the Sn bearing pegmatites of New Ross, Nova Scotia, (2) mica-apatite-pyroxenite pegmatites of Ontario and Quebec, (3) Li pegmatites of southeastern Manitoba and (4) granite pegmatites of British Columbia.

A. M. BRANT

A remarkable occurrence of thucolite and oil in a pegmatite dike, Parry Sound district, Ontario. HUGH S. SPENCE *Am Mineral* 15, 499-520 (1930), cf. C. A. 23, 790.—The location is the Besser mine in Henry township. The pegmatite is a light-pink perthite. The mineral occurred near the dike walls, at or near the surface. It was closely associated with radiated groups of a highly altered mineral now chlorite, thought to have originally been pyrrhotite. The mineral thucolite occurred in part as well-defined cubic crystals, apparently pseudomorphic after uraninite. Massive material is jet black with a conchoidal fracture and a sooty brownish black streak. Its hardness = 4 and sp gr 1.57. The cubes have a dull earthy appearance. Analysis gave

	Thucolite Crystals	Massive Thucolite
Moisture, 110°	2.00	1.60
Volatile hydrocarbons	20.19	19.95
Fixed C	50.62	61.56
Ash	26.86	16.63
	99.87%	99.75%

The ash from both samples was strongly radioactive. Minerals found closely associated with the thucolite were allanite, *beryl*, *biotite*, *cyrtoilite*, *garnet*, *marble*, *pyrrhotite* (?), *titania*, *uraninite*. An oily substance was found in 2 parallel fracture zones across the dike. The occurrence of the thucolite indicates that it is not a primary constituent of the pegmatite, and S. suggests that its origin may have been from oil introduced in a later stage.

A. M. BRANT

A note on the occurrence of monazite in western Arizona. ROBERT E. S. HERNIMAN *Am Mineral* 15, 536-7 (1930).—Scattered pebbles were found in the stream gravels in Mohave county. Some were good crystals, elongated on the b -axis. The sp gr of a group of 8 crystals was 5.04.

A. M. BRANT

The structure of *apatite*. L. MAXIM *Mineral. Z. Krist.* 75, 323-31 (1930).—The dimensions of the hexagonal unit of *apatite* are $a = 9.30$ and $c = 6.85 \text{ \AA}$. $U (=0.02)$. The space group is C_{6h}^2 , and there are 2 mols of $\text{Ca}_5(\text{F}(\text{PO}_3)_4)_2$ in the cell. General positions for all atoms are given, but no parameters. ST. NARAY-SIABO *Ibid.* 387-98 (1930) (in English).—A complete defn. has been made by means of abs. intensity measurements. Parameters are given for all at. positions. There are 6 (PO_3) groups in the unit cell, 4 Ca ions are on trigonal axes, surrounded prismatically by 6 O ions, 6 Ca ions are surrounded by an irregular polyhedron of 1F and 5O. F ions are at the corners of the reflection planes, touching 3 Ca ions each.

L. S. R.

Laralite from Chittenden, Vermont. CHARLES PALACHE and F. A. GONYER. *Am*

Mineral 15, 338-9(1930)—The specimen, deep blue in color, is backed by quartzite containing sericite, rutile needles and hematite grains. Soln of sample for analysis gave 1.49% insol, contg SiO_2 and probably TiO_2 . The sol material had the compn: Al_2O_3 33.11, FeO 2.59, MgO 12.38, H_2O^+ 6.24, P_2O_5 46.17, sum 100.49%; formula $(\text{Mg}, \text{Fe})\text{O} \cdot \text{Al}_2\text{O}_3 \cdot \text{P}_2\text{O}_5 \cdot \text{H}_2\text{O}$, with $\text{Mg} : \text{Fe} = 8 : 1$. The sp gr is 3.08. The optical properties are biaxial— with $2V = 70^\circ$, $\alpha = 1.612$ colorless, $\beta = 1.634$ light blue, $\gamma = 1.643$ blue. A. M. BRANT

Two phosphates from Dehrn, dehrnite and crandallite. ESPER S. LARSEN AND EARL V. SHANNON *Am Mineral* 15, 303-6(1930)—The 2 minerals were found on a specimen of brecciated phosphorite. *Dehrnite* formed grayish or greenish white crusts of fibrous to bladed crystals. Its phys properties are hardness 5, sp gr 3.04, fusibility 2, with formation of a white opaque bead. Basal cleavage, hexagonal outline. Optically it is uniaxial, $\omega = 1.622$, $\epsilon = 1.614$. The compn is: insol 0.12, CuO 50.88, P_2O_5 37.12, K_2O 1.20, Na_2O 7.11, H_2O^+ 1.52, H_2O^- 0.16, CO_2 1.49, sum 99.60%, probable formula $7\text{CaO} \cdot \text{Na}_2\text{O} \cdot 2\text{P}_2\text{O}_5 \cdot \text{H}_2\text{O}$. *Crandallite* occurs in creamy-white radiating prismatic fibers attached to dehrnite. It is optically +, has a moderate optical angle and has $\alpha = 1.59$, $\gamma = 1.60$. X is parallel to elongation and Z perpendicular to the cleavage (100). It is probably orthorhombic. The compn is: SiO_2 4.92, Al_2O_3 37.52, P_2O_5 25.24, CaO 11.04, MgO 0.24, H_2O^+ 17.00, H_2O^- 1.00, CO_2 2.54, sum 100.40%, formula $\text{CaO} \cdot 2\text{Al}_2\text{O}_3 \cdot \text{P}_2\text{O}_5 \cdot 6\text{H}_2\text{O}$. This mineral has previously been called "kalkwavelit". A. M. BRANT

The minerals of the phosphate nodules from near Fairfield, Utah. ESPER S. LARSEN AND EARL V. SHANNON *Am Mineral* 15, 307-37(1930)—The nodules were found embedded in a matrix of quartz and calcite. They have an ellipsoidal or discoidal shape and vary in size up to 8 in. The methods used in the study and analysis of the minerals are described. Most of the nodules consisted of a no. of minerals arranged in concentric layers or in part lining or filling cavities. Calcite, variscite, pseudowavellite and wardite were present in most of the nodules. *Dehrnite* (cf. preceding abstr.) was observed in cavities of the pseudowavellite. Seven new minerals were segregated and studied optically and physically and analyzed. The results are tabulated. *Wardite* was found to be identical with soumanite. *Pseudowavellite* is the most abundant mineral in the nodules. The new minerals are *delavite*, $8\text{CaO} \cdot 5\text{Al}_2\text{O}_3 \cdot 4\text{P}_2\text{O}_5 \cdot 14\text{H}_2\text{O}$, *dennisonite*, $6\text{CaO} \cdot \text{Al}_2\text{O}_3 \cdot 2\text{P}_2\text{O}_5 \cdot 5\text{H}_2\text{O}$, *levistonite*, $15\text{CaO} \cdot (\text{K}, \text{Na})_2\text{O} \cdot 4\text{P}_2\text{O}_5 \cdot 8\text{H}_2\text{O}$, *englishite*, $4\text{CaO} \cdot \text{K}_2\text{O} \cdot 4\text{Al}_2\text{O}_3 \cdot 4\text{P}_2\text{O}_5 \cdot 14\text{H}_2\text{O}$; *millsite*, $2\text{CaO} \cdot \text{Na}_2\text{O} \cdot 6\text{Al}_2\text{O}_3 \cdot 4\text{P}_2\text{O}_5 \cdot 17\text{H}_2\text{O}$, *lehute*, $6\text{CaO} \cdot (\text{Na}, \text{K})_2\text{O} \cdot 4\text{Al}_2\text{O}_3 \cdot 4\text{P}_2\text{O}_5 \cdot 12\text{H}_2\text{O}$, *gordonite*, $\text{MgO} \cdot \text{Al}_2\text{O}_3 \cdot \text{P}_2\text{O}_5 \cdot 9\text{H}_2\text{O}$. A. M. B.

Scorodite from Gold Hill, Tooele Co., Utah. WM. F. FOSHAG, HARRY BERMAN AND RUTH A. DOGGETT *Am Mineral* 15, 390-1(1930)—The color of the crystals is pea-green. The mineral is optically—; by the Merwin dispersion method the ns for D are: $\alpha = 1.784$, $\beta = 1.796$, $\gamma = 1.814$, dispersion $F-C = 0.3 \approx 0.05$. $2V = 54^\circ \approx 5^\circ$. The mineral is a normal scorodite and at Gold Hill forms bodies of considerable magnitude as an alteration product of arsenopyrite. A. M. BRANT

Psittacinite from the Higgins mine, Bisbee, Arizona. STEPHEN TABER AND WALDEMAR T. SCHALLER *Am Mineral* 15, 575-9(1930)—The mineral occurs in coarsely granular masses of anhedral crystals up to 2 cm diam with well-defined cleavage in one direction. It is intimately assoc. with higginsite and platy white barite. The compn is: V_2O_5 21.11, As_2O_3 trace, PbO 50.13, CuO 19.10, ZnO none, ignition loss (H_2O) 4.79, insol 3.06 (MnO_2), sum 98.19%, approximating the formula $2\text{PbO} \cdot 2\text{CuO} \cdot \text{V}_2\text{O}_5 \cdot 2\text{H}_2\text{O}$ ascribed to psittacinite. The color is dark olive-green, streak yellow. The cleavages show parallel extinction, and the mineral is probably orthorhombic. Approx values for n are $\alpha = 2.22$, $\gamma = 2.33$. The dispersion is strong, $p > v$. Thick fragments show distinct pleochroism from yellow to olive-green. The analysis of the mineral from the Higgins mine is compared to that from the Shattuck mine nearby, and the relation to descloizite is discussed. A. M. BRANT

Seamanite, a new manganese phosphoborate from Iron County, Michigan. E. H. KRAUS, W. A. SEAMAN AND C. B. SLAWSON *Am Mineral* 15, 220-5(1930)—The mineral occurs in the Chicago mine near Iron River, Mich., as small acicular crystals in crevices in a highly ferruginous siliceous rock. A fibrous asbestos-like mineral is assoc. with it. The crystals are transparent, pale yellow, unit prisms terminated by pyramids, with distinct basal cleavage, orthorhombic—bipyramidal, with $a : b : c = 0.5195 : 1.04508$, hardness 4, sp gr 3.128. It is optically + with $\alpha = 1.640$, $\beta = 1.663$, $\gamma = 1.665$; $X = a$, $Y = b$, $Z = c$ and $2V$ about 40° . Analysis gave: MnO 56.22, $(\text{Mg}, \text{Ca})\text{O}$ 1.61, FeO 0.13, B_2O_3 9.95, P_2O_5 16.65, H_2O 14.57%. The mineral is regarded as a double salt of $3\text{MnO} \cdot \text{B}_2\text{O}_3 \cdot 3\text{H}_2\text{O}$ and $3\text{MnO} \cdot \text{P}_2\text{O}_5 \cdot 3\text{H}_2\text{O}$ in equal proportions. A. M. BRANT

Crocoite, lautarite and dietzeite and their crystallographic relationships. B. GOSSNER AND F. MÜLLER. *Z. Krist.* 75, 410-20 (1930).—Lattice constants were detd. as follows: crocoite, $a = 6.82$, $b = 7.49$, $c = 7.16$ A. U., $\beta = 102^\circ 33'$, C_{2h}^5 , 4 mols. $PbCrO_4$ in cell, lautarite, $a = 7.18$, $b = 11.38$, $c = 7.42$ A. U., $\beta = 100^\circ 22'$, C_{2h}^5 , 4 mols. Ca_2FeO_4 , dietzeite, $a = 10.16$, $b = 7.30$, $c = 14.03$ A. U., $\beta = 109^\circ 32'$, 4 mols. The compn. is considered as $Ca_2FeO_4 \cdot CaCrO_4$ rather than $7(Ca_2FeO_4) \cdot 8(CaCrO_4)$ as usually stated. The structural and crystallographic relations of these minerals are discussed. L. S. RAMSDELL.

Transparency of gypsum and mica in the ultra-violet. W. HAUSMANN AND O. KUMPEL. *Strahlenkerap.* 35, 387-90 (1930).—Gypsum is highly transparent for ultra-violet rays, which are largely absorbed by mica. B. C. A.

The chemical composition and genesis of the natural iron sulfates. XIV. RUDOLPH SCHAEFER. *Z. Krist.* 75, 67-87 (1930), cf. C. A. 22, 47.—A detailed discussion of the genesis of *romerite*. If free H_2SO_4 is present in an aq. soln. of rhomboclase, hydrolysis is prevented and the acid $H_2Fe(OH)_2(SO_4)_2 \cdot (H_2O)_2$ is stable. This reacts with $(HO)Fe \cdot (H_2O)_2$ to form *romerite*. A series of analyses of natural and artificial voltaite shows it to be a mixt. of the KFe^{++} salts of $H_2HFe(SO_4)_2 + 13Aq$ and $H_2Fe(SO_4)_2 + 18Aq$. L. S. RAMSDELL.

Replacement of wolframite by scheelite, with observations on the fluorescence of certain tungsten minerals. FRANK R. VAN HORN. *Am. Mineral.* 15, 461-9 (1930).—A specimen from the East Pool mine, Cornborne, Cornwall, England, contains wolframite which is partly or wholly altered or replaced by pinkish scheelite. The alteration seems to have taken place at right angles to the cleavage planes. The scheelite shows a bright blue fluorescence under ultra violet rays, of a darker shade of blue than other specimens examd. The results of tests made on a series of W minerals are tabulated. It is suggested that the fluorescence of some wolframites and hübnerites is caused by a slight alteration to scheelite. A. M. BRANT.

A new iron meteorite from Carbo, Mexico. CHARLES PALACIO AND P. A. GONZALEZ. *Am. Mineral.* 15, 383-9 (1930).—The meteorite, found in 1923, weighed about 1000 lb. The surface is covered with the usual rounded depressions, and in addn. there are a no. of cylindrical holes which are thought to have contained troilite. An etched section shows well marked octahedral structure. There are numerous nodules of troilite. The bands of kamacite are regular and from 0.5 to 1 mm. thick. The meteorite is classed as a medium octahedrite, its compn. is given. A. M. BRANT.

Statistical review of Ontario's mineral industry in 1929. W. R. ROGERS AND A. C. YOUNG. *Ann. Rept. Ont. Dept. Mines* 39, Pt. 1, 1-63 (1930). A. H. B.

Quinquennial review of the mineral production of India for the years 1924 to 1928. E. H. PASCOE. *Records Geol. Survey India* 64, 89-92, 251-7, 318-20, 322-4, 350-3, 384, 385-92, 400-3 (1930), cf. C. A. 24, 4245. A. H. B.

The copper shale of Mansfeld. J. W. GIBBOY. *Bull. Inst. Mining and Met.* No. 316, 1-22 (1931), cf. C. A. 25, 898.—Discussion. ALBAN H. EMEY.

Origin of Copper Mountain (British Columbia) ores. V. DOLMACHE. *Trans. Can. Inst. Mining Met.* 32, 151-62.—An analysis is given of the ore, which consists of basaltic and andesitic breccia very little altered by mineralization. Possible modes of origin of the ore are discussed; it was probably formed before the magma had solidified to any great extent. B. C. A.

Copper on the Coppermine River, N. W. T. GEOFFREY GILBERT. *Econ. Geol.* 26, 96-108 (1931).—This region in the past has been very inaccessible, but the airplane has revolutionized these conditions. Native Cu, although of widespread occurrence, gives little promise of economic importance, and the chief Cu mineral is the sulfide, chalcocite. Mineralization is apparently confined strictly to the fissures. Resemblances to the deposits of Michigan are obvious, the main difference being in the angle of dip of the flows, which may have a definite bearing on the question of ore deposition. There appears no reason to suppose that the Cu mineralization will be co-extensive with the basalts on the one hand or strictly limited to them on the other. In this district, as in any other, the main requisite is an adequate supply of Cu bearing solus. Without these, the basalts are no more likely to contain important ore than any other rock, given these, the deposits may occur in other formations as well as in the basalts. ALICE W. EPPASOY.

A quantitative and qualitative determination of the ores of Cobalt, Ontario. JACOB W. YOUNG. *Econ. Geol.* 26, 112-8 (1931).—Y. discusses the recent paper by Ellis Thomson (C. A. 25, 1466). T's conclusions as to the pptn. of Ag in the veins at Cobalt, and also for the time of the deposition of the calcite composing the veins in relation to the

metallic arsenides and sulfarsenides, are at variance with previously formulated views which have much to support them. Y. gives data to support these contentions.

Determination of carbon content in carbonaceous pyrite. V P ZEMLYANTZEV. *J Chem Ind (Russia)* 6, 889-90(1929) —Detns by 3 different methods were made on a closely bound mixt. of coal and pyrite. All gave practically the same figure of 26.3% C.

Origin of the Manchurian magnesite deposit. HIRONAO NISHITARA. *Econ Geol* 26, 118-21(1931) —N discusses a statement made by T Kato (*Econ Geol* 24, 90-3) as to the origin of the Manchurian magnesite deposit. The conclusion as to whether this deposit is a product of replacement of calcareous rock by Mfg of magmatic origin, as Nunomy and Kato conclude, or of sedimentary origin, as N believes, may have an important economic bearing on future exploration for similar magnesite deposits elsewhere.

The witherite deposit of the Settlingstones mine, Northumberland. G TRESTRAIL. *Bull Inst Mining and Met* No 316, 23-6(1931), cf *C A* 25, 899 —Discussion.

A gold and diamond field in Brazil. ANON. *Mining J* 170, 744-5(1930)

Diatomite. PAUL HATMAKER. *Bur Mines Information Circ* 6391, 20 pp (1931). —H discusses diatomite under the headings varieties, occurrence, distribution, properties, uses, mining, prepn, marketing, producers, production, economic condition of the industry and a bibliography of 35 references.

Gypsum. P P BUDNIKOV. *Bull acad sci U R. S. S* 1930, No 77, 1-180 —A collection of the papers published on the above subject by B. and co-workers from 1922 to 1928.

Salt and gypsum in Alberta. J A. ALLAN. *Trans Can Inst Mining Met* 32, 232-54(1929) —An account is given of the distribution of salt and gypsum in Alberta. They are mostly of Devonian or Silurian age. Interbedding of salt, anhydrite and dolomite in the McMurray deposits shows that they have been deposited at the edge of a sea of varying conen. The anhydrite is of primary origin, but much of the gypsum occurring as selenite, is secondary.

The microbiological aspect of peat formation. A. C. THAYSEN. *Fuel in Science & Practice* 9, 560-3(1930) —Analysis of the behavior of all known types of cellulose-decomposing microorganisms is given as proof that their activity is limited to the very surface layers of peat, that it decreases in intensity under an excess of moisture and that in acid peats it ceases long before the cellulose has been eliminated.

The present nature of coals resulting from their history. LEON DUBRUL. *Rev. univ mines* 5, 10-5, 40-7, 65-77(1931) —It is difficult to draw clear-cut conclusions as to the effect of the history of coals on their nature. The single conclusion is that the causes of differentiation are many. The evolution of the coal beds is as follows: The peat formed had a variable chem. compn. depending on the nature of the plants and especially on their manner of decompn. and consequently on the rate at which the peat bogs sank. According to the wt. of the overlying material the mass underwent an anthracitization, the modality of which depended upon the initial state, this mass, then under equal pressure and time, showed what is termed the "original differentiation," consisting in a localization of varieties richer in volatile matter in the zone where the subsidence was quickest. But if the overburden varied in thickness, it could have modified the differentiation and at times might have had the reverse effect (Card). At the same time, under the influence of the variable load there was produced a vertical change by a decrease in volume, relatively greater for lower beds (Hilt's law). At the time of folding the influence of static pressure became negligible as compared with dynamic pressure. Orogenic forces were the cause of a new lateral differentiation, because the quantity of energy transformed into heat depended upon the resistance to the lateral thrusts. It was thus that the regions displaced did not become thinner, the rubbing having been solely a source of heat (Serang); the masses that had been crushed against a resistant block for the most part lost energy, from folding, in heat and underwent an intense devolatilization (Herstal). The depth of the bed had only a min influence on the coals already formed. Four differentiations are: (1) original; (2) by wt. before folding; (3) through orogenic forces, and (4) by wt. after folding. The first and third are generally the most important; the second is found occasionally, and the last relatively seldom.

Igneous intrusions and ore deposits of the zone of flowage. E. C. ANDREWS. *Econ. Geol.* 26, 1-23(1931). —The igneous rocks of the Willyma or Broken Hill "older-

mass" belong to 2 periods of intrusion: that of marked folding, and that of relative stability. The ore deposits are phases of igneous intrusions belonging to the closing stages of the "bedded" form. All the ore deposits examined appear to end vertically as well as laterally, either against or between bedding planes. No sign of feeding source has been detected for the igneous rocks. A. infers that the Broken Hill region during the earlier portion of pre Cambrian time was affected by a movement of strong folding in which there was a strong reaction toward the surface of the earth. Field study suggests that the igneous rocks and the ore deposits may not contain the same mineral assemblages which they possessed during their passage to the "pressure-slacks." The latest stages of igneous intrusions appear to have been attended with increased mobility of the igneous material, resulting in injection and replacement, both of sediments and the earlier orthogneisses, by aphte, pegmatite and ore deposits. A. W. E.

Structural geology of the Conception Bay region and of the Wabana iron ore deposits of Newfoundland. ALMAAR O. HAYES. *Econ. Geol.* 26, 44-64 (1931).—A table of formations is given, and stratigraphy and structural geology of the region is discussed. The limits of the Fe ore are fixed by faults, and the northern limits of the ore are thought to be along a fault line from Colliers Bay to the middle of Conception Bay. The Fe ore reserves are discussed, and it is stated that mining operations may continue on a scale limited only by the amt. the market will absorb and the engineering difficulties of submarine operations.

ALICE W. EFFERSON

Southern contact of the Riesen Mountain granite and the adjacent slate region. ADOLPH WATZNAUER. *Lotus* 78, No. 2, 112-G1 (1930).—A petrographical study.

W. L. HILL

Differentiation in the Cape Spencer flow. RICHARD J. LUND. *Am. Mineral.* 15, 539-65 (1930).—This is the thickest and lowest of a series of 3 Triassic lava flows found on Cape d'Or, Nova Scotia. The optical study was made on material from a drill core. The variations in mineral content, in the FeO/MgO ratio, and in the optical properties of the pyroxene and plagioclase series are given in the form of tables and graphs. Differentiation of the type known as fractionation took place during the crystallization of the flow. This was brought about by the fractional crystallization of the plagioclase and pyroxene series together with the relative movement of these minerals with respect to the liquid.

A. M. BRANT

Differentiation in the Dartmoor granite. A. BRAMMALL. *Nature* 126, 132 (1930).—Graded and significant variations in the chem. compn. of granites from East Dartmoor are discussed.

B. C. A.

Eruptive rocks of the Oberwiesenthal, Erzgebirge. R. HERRE. *Chem. Erde* 4, 632-65 (1930).—Petrographical descriptions with chem. analyses are given of phonolites and basalts from this locality in Saxony.

B. C. A.

Geology of the south coast of New South Wales. III. The monzonitic complex of the Mount Dromedary district. IBA A. BROWN. *Proc. Linn. Soc. N. S. Wales* 55, 637-98 (1930).—Tables of complete chem. analyses of many specimens of the rocks are given. The igneous complex is intrusive into metamorphosed sediments of early Paleozoic age, the main intrusion being an outcrop over an oval area of about 25 sq. mi., probably in laccolithic form. The igneous complex consists of plutonic and hypabyssal types. A study of the field occurrence and assocns. of these igneous rocks and of their mineralogical and chemical compns. shows that probably they are all comagmatic, but that the degrees of consanguinity of the different groups vary. It considers that the main monzonitic series composing the laccolith has been produced as a result of the differentiation of a monzonitic magma in place, by means of fractional crystallization and the sinking of crystals. The origin of the garnet bearing series is more obscure, it may be due to secondary differentiation of a basic phase of the magma, or the result of assimilation of limestone by the monzonitic magma, with subsequent differentiation. Direct evidence of the age of the intrusion is scanty, but the chem. similarity of the series as a whole to the monzonitic rocks of Milton and the latitic flows of the Illawarra District indicates a probable Permian-Carboniferous age. ALICE W. EFFERSON

The iron-rich norite at Akkavare (Sjaunja). PER GEIJER. *Geol. Foren. Forh.* 52, 391-7 (1930).—Varieties of norite rich in Fe have been discovered at Akkavare, Sjaunja. The rock is generally massive, dark and with the size of the grain averaging 1 to 2 mm. for the main constituents. Five specimens are described in detail. The magnetite and ilmenite content varies from 10.7 to 28.8% by vol. in the different specimens.

WILHELM SEGERBLOM

Lamprophyrite of the petrographical province of Predazzo. SILVIO VARDABASSO. *Ann. scuola mag. Padova* 4, 195-216 (1928).—In the eruptive territory of Predazzo and

Monzone, the stratigraphical series and the volcanic rocks are crossed by numerous dark veins. These may be put into 2 classes: (a) non-differentiated veins related to porphyrite and melaphyre traversing all sedimentary rocks remaining in the district (Triassic period), tuffs and lava but not the intrusive masses, (b) veins differentiated by the continuation of intrusive masses (monzonite, etc.), these are the most recent of all rocks in this part of the Dolomites. The latter basic series, or lamprophyrites, contain in turn many types considered as representing an alk-calcareous magma (pacific) and others an alk. magma. This is one of the most salient characteristics of this petrographical province. V gives 4 figures illustrating the compn. of camptonite, 3 for monchiquite, 4 for rizonite, with a table of compn., and 2 for lamprophyrite with 3 analyses. There is a bibliography of 34 references. S. L. B. ETHERTON

Water content and degree of compactness of argillaceous rocks. W. PRITASCHER AND B. WILSER. *Berg. Huttenm. Jahrb.* 74, 57-65 (1926).—The water content, other than that evolved at a red heat, diminishes with increasing metamorphism. B. C. A.

The mineralogy of some deposits of kaolinized volcanic ash from the slate belt of North Carolina. JASPER L. STUCKEY. *Am. Mineral.* 15, 253-8 (1930).—The deposits described lie in areas of acid volcanic rocks. Quartz, altered feldspar, sericite, kaolinite, epidote, rutile, titanite, hydromica, and halloysite were identified by microscopic means, while much fine-grained material could not be identified. Screen tests and chemical analyses of the material are given. The origin of the deposits was fine-grained tuff or volcanic ash. In the northern part they have been changed by severe metamorphism to sericite schist, and in the south have weathered to impure kaolin. A. M. B.

Mechanical analyses of sediments by centrifuge. PARKER D. TRASK. *Econ. Geol.* 25, 581-99 (1930).—An application of the centrifuge to the procedure of mechanical analyses of sediments on the basis of Stokes' law is described. The method consists in decanting the sand and detg. the rate of increase in wt. of finer particles settling out, with time. This is done by sepg. the suspension contg. these finer constituents into a no. of aliquots and detg. wt. of material that settles from each aliquot after centrifuging for definite times and at specified speeds. Curves representing the rate of fall are constructed by plotting the wt. of sediment that settled from the aliquots against the proper values for the time. From these curves, both the size distribution of the constituents and the percentile diams. are readily obtained. Chief advantages of the method are: it is rapid and does not involve complicated app., it permits of presentation of the results as percentile diams.; and it gives the complete size distribution of the constituents down to any desired dimension, no matter how small. A. W. E.

Sedimentation in the Channel Islands region, California. PARKER D. TRASK. *Econ. Geol.* 26, 24-43 (1931).—The mech. compn. of the deposits of this region varies with the configuration of the ocean bottom. Finer sediments accumulate in the basins and coarser on the divides, topography influencing sediments more than depth of water or distance from shore. Finer deposits of the Pacific Coast region are more coarse-grained than corresponding sediments forming off master streams, such as Orinoco, Mississippi or McKenzie. The CaCO_3 content of sediments is variable, but in general less in shallow water than in depressions, where it reaches a max. of 16%. The org. content varies with bottom configuration and ranges from less than 1 to 75% in deposits in the bottom of Santa Cruz basin. Ignition loss is a rough index of the amt. of org. matter. The C-N ratio of recent deposits does not vary greatly from 8.4. Consequently, the org. content of sediments may be estd. by multiplying the percentage of N by 14. Sediments of past geologic age contain about half as much N with respect to C as do recent deposits. ALICE W. EPPERSON

Subterranean water conditions in the coastal regions of the Netherlands. J. VERSLUYS. *Econ. Geol.* 26, 65-95 (1931).—Geologic history of the area is given, and the character of the deposits and their influence on ground water are described. Chem. action in the subsoil of the dunes is discussed. Formation of water with excess of alkalis is explained by assuming that the sea-water has ceded alkalis to the soil, and that these, when the sea-water was displaced by calcareous fresh water, have in turn been exchanged for Ca. ALICE W. EPPERSON

Iron and manganese hydroxide soils in relation to the black coating on rocks and the formation of laterite. W. KNAUST. *Chem. Erde* 4, 529-48 (1930).—Soils of the hydroxides were prepd. with H_2PO_4 as a peptizing agent, and various expts. were made to det. the conditions for their coagulation. The bearing of these expts. on the deposition of Fe and Mn oxides as a coating on rocks and on the formation of laterite is discussed. B. C. A.

Soils derived from the glacial boulder-marl in the Hamburg district. E. SCHMIDT. *Chem. Erde* 4, 475-500 (1930).—Detus. of the amt. of CaCO_3 , of the size of particles,

and of the heavy minerals present in the boulder clay and the overlying soils are recorded. B. C. A.

Annual review of Swedish geological literature for 1929. R. SANDEGREN and M. MAGNUSSON *Geol. Fören. Förh.* 52, 689-754(1930)—Continuation with 190 titles of the resumés described in C. A. 23, 1598, 24, 1605 WILHELM SÖGERBLOM

The unipolarity of pressed PbS (TERRY) 2. Quinquennial review of the mineral production of India for the years 1924 to 1928 (PASCOR) 9. The evolution of coal (BRAGGS) 21. Method and cost of recovering quicksilver from low-grade ore at the reduction plant of the Sulphur Bank Syndicate, Clearlake, Calif. (BRADLEY) 9. Radioactivity and the thermal history of the earth (HOLMES) 3. Preparation and purification of the triiodides of Sb and As for use in immersion media of high refractive index (FISKE) 6. Structure of silicates (BRAGGS) 6. Relation between the contents of coal in volatile matter and in ash (LE GRAYE) 21.

9—METALLURGY AND METALLOGRAPHY

D. J. DEMOREST, H. W. COLLETT and RICHARD RIMBACH

What the Bureau of Standards does for metallurgy. RICHARD RIMBACH *Metals and Alloys* 2, 65-70(1931) A. J. MONACK

Principal metallurgical products utilized in aeronautical construction. G. R. HAMEL *Aviers* 5, 553-8(1930) A. J. MONACK

Progress of metallurgy at Broken Hill, N. S. W. M. R. McKNOWLY *Chem. Eng. Mining Rev.* 23, 93-100(1930)—The equipment, operation and reagents used are described. Slow speed crushers and vibrating screens are preferred. An all flotation treatment with an alk. current of 0.005% Na_2CO_3 is used. This process yields higher-grade Pb and Zn concentrates, low tail factor for flotation and lower costs. A. N. H.

Quinquennial review of the mineral production of India for the years 1924 to 1928. E. H. PASCOR *Records Geol. Survey India* 64 (1930)—Introduction. *Ibid* 1-3. Summary of progress. *Ibid* 4-25.—Detailed production figures for Indian minerals from 1924 to 1928 are given. Antimony. *Ibid* 26-8.—Stibnite lodes in gneissose granite occur in Punjab. They carry 6 dwts. of Au per ton. Stibnite with cerussite occurs in the Northern Shan states. The tetrahedrite in the Sleemanabad Cu lodes is highly antimonial. About 1000 tons per yr. of antimonial Pb, analyzing Pb 77, Sb 21% and Ag 6-8 oz. per ton, is produced in Burma. Copper. *Ibid* 75-83.—In Singhbhum, Cu lodes occur in muscovite and chlorite-quartz schists. At the surface they consist of quartz with malachite, chrysocolla and red oxides of Fe contg. small quantities of Cu, possibly as red oxide, with sometimes small incrustations of lebetheite. In depth the ore is entirely chalcopryite. There is a small chalcocite zone not far below the surface. In the Dargelling area, chalcopryite occurs disseminated through slates and schists. Recent work has shown lodes in Sikkim in which Cu is associated with Bi, Sb and Te. Tetradymite and linnaeite are found. Typical minerals are chalcopryite, pyrite, pyrrhotite, sphalerite and galena. Several occurrences are described. Bawdwin (Burma) ore analyzed Pb 13, Zn 8, Cu 7% and Ag 18 oz. per ton. Arsenic. *Ibid* 320-2.—The chief Indian source of As is the orpiment mines of Chitral where realgar, orpiment and fluor spar occur. Bismuth. *Ibid* 344.—Native Ni and bismuthinite occur in wolfram and cassiterite-bearing veins of lower Burma. Cobalt. *Ibid* 379-80.—Cobaltite and danasite have been found disseminated among certain slates. The Ni spess of the Burma Corp. contains 3-4% Co. Nickel. *Ibid* 412-3.—Nickeliferous pyrrhotite, chalcopryite and pyrite occur in various parts of India. It is a by-product in the smelting operations of the Burma Corp. ALDEN H. EMERY

Quinquennial review of the mineral production of India for the years 1924 to 1928. Iron. H. CECIL JONES *Records Geol. Survey India* 64, 111-46(1930)—Bihar and Orissa are the only parts of India where Fe ore is mined for the production of Fe and steel. In Burma limonite and hematite are produced for use as a flux in Pb smelting. The main plants are described and analyses given of their raw materials and finished products. Ore occurrences are also described in detail. ALDEN H. EMERY

This concentration of ores by flotation. H. L. SULMAN. *Bull. Inst. Mining and Met.* No. 316, 31-6(1931), cf. C. A. 24, 2699, 5006, 5263, 25, 901—Discussion

Water concentration tests. BERNARD W. HOLMAN *Bull. Inst. Mining and Met.* No. 316, 29(1931), cf. C. A. 24, 2699, 25, 901—Discussion. ALDEN H. EMERY

Reduction of iron oxide ores in molten state by means of hydrogen and other gases. SAMUEL L. MADORSKY. *Ind Eng Chem*, 23, 99-103 (1931)—Thirteen expts. were undertaken to reduce molten Fe oxide with H_2 , water gas and CO. For each kg of Fe 1.19 cu m. of H_2 , 45% of the theoretical, was required. An Fe product is obtained at a cost comparable to that of other steel-making processes. C. H. LORIO

Leaching copper ores: advantages of wet-charging. JOHN D. SULLIVAN AND ALFRED P. TOWNE. *Bur Mines, Rept of Investigations* 3050, 20 pp (1931), cl. C. A. 25, 478.—Charging the ore into the leaching vat in the wet condition increases the extn of Cu. Virtually identical extn was obtained by upward and downward percolation when the solns. were not continuously circulated. In the latter case, extn. was greater on downward percolation. Virtually identical extn resulted from flood and trickle leaching when the ore was wet charged. With dry-charged ore, extn. was greater from +100-mesh ore than from unsized material contg fines, there was no difference when the ore was wet charged. ALDEN H. EMERY

Cyanidation of a copper-gold ore. E. H. SMITH. *Can Chem Met* 15, 45-6 (1931)—The function of lime in cyanidation of Au ores is attracting much attention. Lime and alkyl have no solvent action on the precious metals, but enter into numerous reactions occurring in the process. A pyritic ore ground in lime soln. yields CaS and $CaSi_2O_6$, and gang constituents consume a certain amt. of lime. Each mineral in an ore removes lime from the mill soln in an amt. dependent on the lime present, and this amt. is sp. for each mineral. Lime decomposes the $Zn(CN)_2$ or double cyanide formed in the pptn. processes. W. H. BOYNTON

Method and cost of recovering quicksilver from low-grade ore at the reduction plant of the Sulphur Bank Syndicate, Clearlake, Calif. WORTHEN BRADLEY. *Bur Mines, Information Circ.* 6429, 17 pp (1931)—The ore deposit consists of grains of cinnabar finely disseminated throughout a gang composed principally of altered basalt with subordinate aints of elemental S (1-15%, av. 2%) and opal. Recovery is as follows: (1) screening (cinnabar mostly in fines) and sorting the low-grade ore to obtain a furnace feed contg 8-10 lb. Hg per ton, (2) treatment in a rotary kiln, (3) pptn. of dust from gas stream, (4) condensing Hg from gas, (5) flotation of low-grade condenser mud and (6) retorting high grade condenser mud and high grade flotation concentrates. Costs per ton of ore treated for April, 1930, are: screening and sorting \$0.835, furnace treatment \$1.709, flotation \$0.387, retorting \$0.177, assaying \$0.107, condensing \$0.534, miscellaneous \$0.359 and total \$4.108. ALDEN H. EMERY

Milling methods and costs at the Homestake Mine, Lead, S. D. ALLEN J. CLARK. *Bur Mines, Information Circ.* 6408, 25 pp (1931)—The ore bodies occur in altered dolomitic limestone contg cumingtonite and chlorite with pyrite, arsenopyrite and pyrrhotite. The Au content is about \$5.00 per ton of ore. Amalgamation recovers 62.8% of the Au (chiefly coarse), cyanidation of sand 20.6%, and cyanidation of slime 9.6% more, a total of 93% recovery. Cost of crushing in mine and at hoists in 1929 was 3.8 cents per ton, milling 27.6 cents and cyanidizing 18.9 cents, a total of 50.3 cents per ton of ore milled. Detailed metallurgical data and costs are given. A. H. E.

The Roan Antelope mill and smelter. A. G. MCGREGOR. *S. African Mining and Eng. J.* 41, pt. 11, 373-4 (1930).

Lead smelting in shaft furnace with zinc-rich slags. KARL PRIOR. *Mettall u Erz* 28, 25-33 (1931)—High Zn content is unfavorable for operation of the shaft furnace. The constitution of shaft-furnace slags in general is discussed, as well as the forms in which Zn is found in the slag. Among these are the oxide, silicates, Zn spinel, Zn ferrites, zincates and sulfides. Operating conditions in the presence of high Zn content are discussed. H. STOEZT

Copper and zinc in cyanidation sulfide-acid precipitation. EDMUND S. LEAVER AND JESSE A. WOOLF. *Bur Mines, Tech. Paper* 404, 63 pp (1931)—All of the common Cu minerals except chrysocolla and chalcophyrite are sufficiently sol. in cyanide soln. to cause excessive reagent loss (2-3 lb. NaCN per lb. Cu) unless the cyanide combined with the Cu is regenerated. The $Na_2S-H_2SO_4$ method regenerates an av. of > 80% of the cyanide from solns. used to treat precious-metal ores contg up to 0.5% cyanide-sol. Cu. Solns. used to treat coprite or metallic Cu give large regeneration. The more complex the Cu mineral treated, the less cyanide will be regenerated from the resulting soln. The more nearly all the cyanide is combined with Cu and the less the cyanide is uncombined (free) the greater will be the regeneration. The amt. of Na_2S used varies with the amt. of metals pptd. The acid consumption varies with the strength of the original soln. in both cyanide and lime. The Ag may be pptd. separately from the Cu before the soln. is acidified. The CuS must be removed from the acid soln.,

as it redissolves in alk. cyanide soln. Only a small quantity of Au is pptd by the $\text{Na}_2\text{S}\cdot\text{H}_2\text{SO}_4$, most must be removed by Zn as usual. Cu in cyanide soln. up to 10 lb per ton does not diminish the activity of the soln. as a solvent for Au and Ag if the requisite free cyanide is present. The double Na-Cu cyanide soln., in the absence of free cyanide, is only a weak solvent for the precious metals and is more effective for Au than for Ag. Smithsonite, hydrozincite, zincite and calamine are sol. enough under the usual conditions of cyanidation to cause rapid accumulation of Zn in the soln. unless special precautions are taken to remove it. Willemite, sphalerite and franklinite are dissolved more slowly. From 1.5 to 4.0 lb NaCN is needed per lb Zn dissolved. The various Zn cyanide compds. are only weak solvents for the precious metals, even if AgNO_3 titration shows free cyanide. Addn. of excess CaO or caustic soda improves extn. by liberating free cyanide from the double salt. If most of the cyanide present in a soln. is combined with Zn, more free cyanide must be added to this soln. than to a fresh soln. for the 2 to have = activity as solvents. $\text{Na}_2\text{S}\cdot\text{H}_2\text{SO}_4$ treatment will remove Zn and regenerate cyanide. When ores contg. sol Zn minerals are treated, the soln. must contain free cyanide, or the soln. of Zn will decrease Ag extn. The Welch method was thoroughly tried and seems to apply only to Au ores or to Au-Ag ores contg. not over 0.1% of cyanide-sol Cu and in which Ag occurs as a haloid mineral. A. H. E.

Recovery of zinc from ferrite compounds in the electrolytic zinc process. G. L. OLDRIDGE and D. P. NICHOFF. Dept. Mining and Met. Research, Univ. of Utah, Tech. Paper 6, 45 pp (1929).—Zn ferrite formation increased with increasing time, temp. and fineness. Fe in Zn ferrite was more readily sol. than when present in Fe_2O_3 . Replacement of ZnO in ferrite by other bases such as CaO and MgO is possible but not practical. Zn ferrite was readily decompd. by treatment with reducing agents at elevated temps., Zn was then sol. in NH_3 or H_2SO_4 . High concns. of reducing gases make the Fe sol. in H_2SO_4 . C will reduce ferrites at 800° . Pure CO should not be used below 800° , as CO_2 is needed at low temps. to prevent its decompn. to CO and C. The best procedure was treatment for 2 hrs. at 550° with CO, contg. 35–50% CO. This makes 80% of the Zn and 39% of the Fe sol. in H_2SO_4 . To make more Zn sol., higher proportions of Fe were dissolved. The rate of decompn. of ferrite with reducing agents, as well as its rate of soln. in H_2SO_4 , HCl and HNO_3 , depends upon the temp. and time of prepn. of the ferrite. By giving refractory ferrite a preliminary low-temp. roast, the subsequent rate of reduction can be increased. ALBERT H. EMMETT

Purifying zinc by dry method. D. M. CHIRZIKOV. *Mineralnoe Suro'e i Tselnost'* Met. 4, 1211–8 (1929).—The processes of purifying Zn are described briefly. M. A. J.

Memorandum on calcination of zinc carbonate ores in particular in a continuous-charged furnace (Jalabert method). GUSTAVE PIXON. *Rev. ind. minérale* No. 241, 9–14 (1931).—By eliminating the moisture and CO_2 , the wt. is decreased 25–32%, according to the ore. The beneficiation of the ore increases its salability. However, calcined ore has to be kept covered in transportation and in storage. Lean calamines are calcined either in an ordinary reverberatory furnace or in an Oxland or Spirek furnace. Best results from calcining are obtained in a shaft furnace, but lean ores cannot be treated in this furnace. The Jalabert furnace, designed to treat this class of ores with the advantages of the shaft furnace, is constructed of brick; ordinary bricks are used as high as the inlet air holes, and the upper part is made of a shell of refractory brick. Ventilation is induced by a blower. In calcining, the lean calamines are mixed first with a combustible in a proportion detd. by lab. tests. At Gustar, wood-charcoal dust is used in the proportion of 9% by wt. of the raw ore. The yield of calcined material varies from 63 to 73%. Calcination of Zn carbonate in the Jalabert furnace takes place under relatively the same conditions as in a shaft furnace and has the following advantages: (1) good utilization of inferior quality combustible that consequently reduces costs; (2) inexpensive construction and installation; (3) reduced labor; (4) continuous operation; and (5) perfect calcination. C. W. OWINGS

A consideration of some of the problems of the chloride volatilization process. ROBERT H. BRADFORD and CAROL MAX MACFARLANE. Dept. Mining and Met. Research, Univ. of Utah, Tech. Paper 3, 21 pp (1923).— CaCO_3 was increasingly deleterious to chloride volatilization as its content in the gang of an ore increased above 20%. Water vapor was beneficial when treating ores with a high calcareous gang. HCl was a more active chloridizer than CaCl_2 or NaCl. HCl was prepd. by spraying a soln. of CaCl_2 or NaCl into the roasting furnace, SiO_2 being necessary with NaCl. An oxidizing atm. contg. 4–11% O_2 was best when chloridizing a Cu ore. A. H. E.

Recovery of noble metals from waste. A. IERVING. *Acta Univ. Lathriensis, Kim. Fakult. Seriya 1*, No. 1, Burtseva 1–11 (in German 12) (1929).—Smelting gives better results than the Cl or CN processes for 1–3% content of noble metal, the efficiency

increasing with the richer charges. The yield is increased by adding CaCO_3 and MgCO_3 and holding for sufficient time at 1350–1400°. ANN NICHOLSON *Ibid.*

Platinum. PAUL M. TYLER AND R. M. SANTMYERS. *Bur. Mines, Information* 6389, 69 pp (1931).—The following subjects are discussed for Pt: properties, uses, alloys, substitutes, history, ores and minerals, occurrence (U. S. and foreign, by states and countries), assaying, mining, concg., refining, production, marketing and prices. *Osmiridium* is briefly discussed. ALDEN H. EMERY

Fluxes for brass melting. E. T. RICHARDS. *Gieseler-Zig.* 27, 433–5 (1930).—Fluxes used in brass melting serve 3 purposes: they accelerate the melting process, they protect the metal from oxidation, and they remove impurities. Glass and sand mixts. are the common com. fluxes, used often with covers of charcoal, coke and NaCl . In crucible melting, borax is used to dissolve SiO_2 , Al_2O_3 , CaO , etc., and to lessen the Zn loss. Mixts. of boric acid, borax and the alk. carbonates are strongly desulfurizing. Boric acid is gradually displacing borax as a flux, in spite of its higher cost. Potash and soda are used to remove S and As. KHSO_4 is valuable for removing Al, but it introduces a little S into the brass. For the removal of Fe, the brass is treated with a mixt. of 75% calcined soda, 18% KCN and 7% powd. glass. For Pb-rich alloys, mixts. of talc and quartz are recommended. CURTIS L. WILSON

Manufacture of wire bars from secondary copper. P. SIEBE. *Gieseler-Zig.* 27, 505–9 (1930), cf *C. A.* 23, 4915. CURTIS L. WILSON

The production of high-test cast iron. H. RICHARD MOLDENKE. *Iron & Steel Can.* 14, 27–8, 41–2 (1931), cf *C. A.* 25, 1192. E. H.

Processing 3-ton wrought-iron balls. A new departure. J. D. KNOX. *Steel* 87, No. 20, 43–6 (1930).—The Aston process for making wrought iron and the new plant of the A. M. Byers Co. near Pittsburgh, Pa., are described. LESLIE B. BRAGG

The reductivity of over-trusted coke and the production of a carbon-poor forgeable channel iron in the foundry shaft furnace. E. PIKOWSKY. *Die Gieseler* 17, 1149–52 (1930).—Expts. to det. the effect of impregnating coke with milk of lime on its reactivity. Thus, a current of CO_2 is allowed to react with coke at 900° and 1100°, and the amt. of CO formed is detd. These expts. show that when coke is satd. with milk of lime CO is formed in greater amt., the reductivity of the coke being actually favored by the impregnation. Claims in the patent of June 2, 1919 (class 18a, Gr. 3, C. 38 289) are disproved. Low-C Fe (1.6% C and less) has successfully been prepd. by P. in the cupola furnace. J. BALOZIAN

Agglomeration of blast-furnace dust and treatment of iron carbonate minerals by the Greenswail process. ROBERT CORDONNIER. *Rev. Metall.* 27, 467–78 (1930).—A description of plant operations and the process for sintering blast-furnace dust and Fe carbonate ores are given. C. H. LORIG

Deoxidation of steel with silicon. C. H. HEATY, JR., G. R. FITZGERALD AND C. F. CHRISTOPHER. *Bur. Mines, Tech. Paper* 492, 42 pp (1931).—The liquidus of the system FeO-SiO_2 has been studied over the range 45–100% FeO. The m. ps. of synthetic Fe^{++} silicates were detd. with the micropyrometer. Fayalite (2FeO-SiO_2) melted at 1335°, two eutectics, at 65 and 78% FeO, were found, their m. ps. were 1260° and 1240°, resp. The deoxidation const. for the reaction $2\text{FeO} + \text{Si} \rightleftharpoons \text{SiO}_2 + 2\text{Fe}$ was detd. to be $K_p = (\text{Si})/(\text{FeO})^2 = 1.65 \times 10^{-4}$ when all concns. are expressed as wt. %. Fluxing of SiO_2 by dissolved FeO was shown to be an important factor in the deoxidation of low-C steel. Steels high in SiO_2 are red-short in forging, while steels contg. Fe^{++} silicates high in FeO forged without cracking. The use of Si as a deoxidizer in steel making processes always results in the formation of SiO_2 or silicates, depending on the amt. of Si used and the concurrent use of other deoxidizers. If Si is used alone, the SiO_2 formed is extremely hard to eliminate. The oxidation of the metal in the basic open-hearth process increases as the C content decreases. The time required to eliminate silicates, resulting from deoxidation with Si, from the steel bath was proportional to the depth of the bath and inversely proportional to the size of the particles formed on deoxidation. Agitation of the bath assisted in cleaning it. High residual Mn had only a slight cleaning effect. Segregation of silicates in ingots from 20 lb. to 3 tons in wt. was studied. In top-poured steels the max. segregation was in the center of the ingot about 15% from the bottom. An annular zone of high-silicate segregation was found in the upper portion of these ingots and is due to crystn. phenomena. In bottom-poured ingots segregation is more irregular and is complicated by erosion of runner brick. ALDEN H. EMERY

Operation and metallurgy of a 200-ton tilting furnace for the Talbot process (of steel making). W. ALBERS. *Stahl u. Eisen* 51, 117–28 (1931).—In continental practice, open hearth furnaces for steel making are considered economical up to 75–100 tons

capacity. For larger sizes tilting furnaces are more economical because of the easy removal of the slag by tilting the furnace. The construction and the operation of a 200 ton tilting furnace, which can be used for either the scrap iron or the pig iron-ore (Talbot) process, is described in detail. In using the scrap iron process best results were obtained by tapping off $\frac{1}{2}$ of the content of the furnace at the end of the melt, while in using the Talbot process the most economical operation was attained by tapping off only 65-75 tons. The chem. reactions during the melting process were followed by taking samples from the melt from various depths. About 3 hrs was required to obtain the steel of the desired compn. For deoxidation Fe-Mn gave better results when added as a liquid rather than as a solid. The mech tests carried out on a great no. of melts showed that the steel obtained by the Talbot process was of a very good quality. Many references are given. J. A. SZILARD

Graphical calculation of cupola furnace mixtures with special regard to pearlitic castings. BÉLA SZÖKE. *Die Giesserei* 17, 1013-7(1930).—A theoretical paper in which is described a method using parallel coordinates for detg. the wt. of each constituent to be charged into a cupola furnace to obtain a given content of a given element (e.g., Ni) in the Fe drawn-off. The charged mixt. is prepd. from 2 or 3 kinds of Fe, with known Si content. Three examples showing the application of the method to various charges (2 pearlitic castings) are given. J. BALOZIAN

The method of running the Buess "rocking revolving" furnace. WILHELM BUSS. *Die Giesserei* 17, 1121-2(1930).—A brief description is given of the Buess furnace for mixing various alloy mixts. by revolving off-center. The furnace is also suited for the production of high-quality cast Fe; it is provided with a device for observing the smelting process and for drawing-off samples without disturbing the melt. When revolving at a speed of $\frac{1}{2}$ revolutions per min., the melt is thoroughly mixed and well degassed, and the thermal efficiency of the fuel is high. The furnace burns either gaseous, liquid or powd. fuels. J. BALOZIAN

Firing with powdered coal in German tempering foundries. RUDOLF STÖTZ. *Die Giesserei* 17, 1112-21(1930).—A lecture. J. BALOZIAN

Structural metallography. H. B. POLSIFER. *Metals and Alloys* 2, 84-8(1931).—Surfacing and etching, etching reagents and metal-fine structure are discussed.

Automatic metallographic polishing machine. W. F. DAVIDSON. *Metals and Alloys* 2, 89-91(1931). A. J. MONACK

Crystallographic investigation of some mechanical properties of metals. I. Y. KIDANI. *J. Faculty Eng. Tokyo* 19, 1-6(1930).—The form of Al crystals after plastic bending has been investigated by means of Laue photographs. II. *Ibid.* 7-15.—Rotational slip occurs in the uniform bending of Al crystal plates. B. C. A.

Kinetic measurements of transformation reactions in solid metals. W. FRAENKEL AND E. WACHSMUTH. *Z. Metallkunde* 22, 162-7(1930).—The decompn. of an alloy of 80% Zn and 20% Al as a result of the transformation occurring at 260° was studied by means of an especially sensitive dilatometer. The velocity of the decompn. is expressed in terms of length (vol.) changes as a function of time. It is found that the decompn. in a quenched alloy near 0° begins at a very low rate, then auto-accelerates and finally comes to completion. The velocity is greatly dependent upon the quenching temp.: the higher this temp., the lower the velocity. An attempt is made to derive an equation expressing the vol. of the reaction. The elec. cond. in quenched alloys first increases and then decreases. The elec. cond. curve over the temp. range 200-360° shows variations at the transformation point, the meaning of which is somewhat uncertain, appearing most nearly like points of inflection. ROBERT F. MEHL

Wear of metals. SAMUEL J. ROSENBERG AND HARRY K. HERSCHEMAN. *Metals and Alloys* 2, 52-6(1931). A. J. MONACK

The effect of cold-working on the density and electrical resistance of metals. TARO UYEDA. *Science Repts. Tôhoku Imp. Univ.* 19, 473-98(1930).—When cold working consists of rolling or drawing, the internal stress cannot be measured. In these expts. the tensile stress applied to test pieces of Swedish steel, Cu and brass was made in an Amsler testing machine. The elec. resistance, elongation and reduction in area were measured at const. temp., and the changes in d and in sp. elec. resistance were calcd. The d decreased in proportion to the stress up to the yield point, after which it decreased more rapidly. The sp. elec. resistance increased in proportion to the stress up to the yield point, after which the Cu and brass showed more rapid change than the steels. The rate of change of sp. resistance decreased with increasing C content. The decrease in d. is attributed to an elastic origin up to the yield point; beyond this point it is attributed to minute cavities formed by internal shippings in the

metal. The increase in elec. resistance is attributed to the same cause above the yield point, below the yield point the increase is explained by the diminution of free electrons, which have become bound electrons around a nucleus caused by enlarging the space lattice.

What is this thing called fatigue? H. W. GILLET *Metals and Alloys* 2, 71-9 (1931).—Review. CURTIS L. WILSON
A. J. MONACK

Some notes on blue brittleness. LELAND R. VAN WERT *Am. Inst. Mining Met. Eng., Tech. Pub. No. 404*, 11 pp (1931).—Stress strain diagrams of low-C steels tested between 200° and 350° depart from the normal regularity of such diagrams secured by testing above or below this range. Under a uniformly increasing load, deformation does not proceed regularly. Periods of deformation alternate with periods of little yielding. This 3 stage cycle continues until fracture ensues. The relative lengths of the periods increase as loading proceeds. It is now believed that aging after overstrain and blue brittleness are the same phenomenon. An effect of aging after overstrain is recovery of the elastic limit. At 250° this aging may produce as complete recovery in a few sec. as in 2 weeks at ordinary temps. Aging after overstrain and blue brittleness are the result of dispersion hardening. High purity Fe gave a decidedly irregular curve at 300°, but without the sharp steplike characteristics usually obtained from steel or Fe contg. more impurities. It is believed that the recovery of elasticity on aging is more apparent than real, and that aging after overstrain affects the elastic properties of ferrite not by actual restoration of the true elastic limit, but by producing through dispersion hardening an effect that simulates recovery. When the apparent recovery becomes instantaneous at blue-heat temps., as with ferrite solns. with usual concns. of dissolved hardeners, or at lower temps. where high concns. result in a greater metastability, the stress-time diagrams take on the steplike character noted.

Directional properties in cold-rolled and annealed copper. ARTHUR PHILLIPS AND E. S. BURN *Am. Inst. Mining Met. Eng., Tech. Pub. No. 413*, 18 pp (1931).—In order to avoid serious directional properties in Cu, it appears desirable to limit the final reduction to 50-60% and the final anneal to 500-600°. It is believed, however, that the preliminary history of the material has considerable bearing on the problem. There is reason to think, for instance, that if the last annealing temp. were 800° instead of 550°, directional properties of a different nature and order would be encountered.

A new etching reagent for nitrified layers. O. HENGSTENBERG AND P. BORNEVELD. *Krupp. Monatsh.* 11, 285-6 (1930).—The new etching method proposed consists first of brief treatment with a soln. of 1 g. HgCl₂ in 20 cc. of concd. HCl and 80 cc. of alc., followed by etching with 10 g. of MgCl₂ and 1 g. of CuCl₂ in 40 cc. of concd. HCl and 100 cc. of alc.

Supplementary discussion of the drop experiments of EVANS. E. MAAS AND E. LIEBREICH. *Korrosion Metallschutz* 6, 172-3 (1930).—This is a reiteration of the belief in the convection currents explanation of the phenomena observed by Evans (cf. *C. A.* 24, 3744, 5705).

The hardness testing of electrodeposits and other thin metallic coatings. HUGH O'NEILL. *Trans. Faraday Soc.* 27, 41-51 (1931).—The paper covers results obtained by indentation and by scratch tests, the same instrument being used as a *micro Brinell tester* and as a scratch hardness tester. The factors involved in the hardness testing of thin metallic coatings are reviewed: (1) *surface finish*—the author finds that for his scratch test described he can, in many instances, secure satisfactory results by the use of 000 emery and slight hand polishing with a soft cloth moistened with a non-scratching metal polish; (2) *thickness of deposit*—a graph is given of the calculated minimum thicknesses of coating which may be accurately tested with various indenting tools loaded so as to produce indentations or scratches of various sizes; (3) *optical difficulties in making measurements* are apt to be great when a shallow indentation of relatively large diameter is being measured; a 0.5 mm. ball is suggested for indentation tests; (4) *internal stress in the coating* is mentioned as a theoretical possibility which may cause a slight opening out of a scratch or indentation; (5) *undue softness of the base metal* may affect the hardness result unless the coating is thick enough to carry all the effects of the deformation; the author has successfully tested an Fe film (0.004 in.) stripped and mounted on glass with Canada balsam as a means of overcoming this difficulty; (6) *local variations in electrodeposits* make it desirable to select the same locality for test on a number of plates. A machine designed by the author (*C. A.* 22, 4433; 23, 5456) utilizing a polished 1 mm. hemispherical diamond has been found suitable for the indentation and scratch testing of coatings down to less than 0.0005 in. in

thickness provided they are not much harder than about 400 Brinell. In both the micro-Brinell and scratch tests, loads up to 1 kg. were used. In the indentation tests the results are expressed as "ball numbers" in kg per sq mm, obtained by dividing the load by the projected area of the indentation (πr^2) instead of the spherical area (as required for the Brinell no.) The use of this "ball no." is recommended as being a more scientific value than the Brinell no. and being considerably easier to calculate than the Brinell no. while not differing greatly from it. A similar "cone no." is used in recording the results of micro-indentation tests using a 120° diamond cone. The scratch hardness no. is given by $P_s = (8L/\pi w^2)$ kg per sq mm (L = total load on diamond, w = width of scratch made by hemispherical diamond). The scratch hardness no., under the conditions used, does not vary greatly with the testing load. An empirical conversion curve for obtaining approximate Brinell nos. from scratch hardness nos. is given. Tables showing the results of applying both the above tests to various metals, and the scratch test to Cd deposits of various thicknesses, to electro-deposits of various metals 0.02 mm thick, and to Ni deposits (on steel) from baths of different compns., are given. It is suggested that a 0.5 mm hemispherical diamond indented on a machine of the above general type would be of still greater utility, as harder deposits could then be tested.

EDWARD B. SANIGAR

New method of making ball hardness tests of metals, and a plotting table for the simplification of Meyer (hardness) analysis. HUGH O'NEILL, *Iron and Steel Inst.* (London), *Carnegie Scholarship Mem.* 19, 19-38(1930).—A crit. discussion of existing hardness-testing machines is followed by a description of a new diamond hardness test which uses an "unrecovered" indentation, dispenses with a measuring microscope, gives comparable hardness nos., has a const. rate of application of load and can be used for standard Brinell testing. The instrument is suitable for Meyer's hardness analysis. For hard materials a 1-mm indenter is employed, while for soft metals and alloys a 10-mm. indenter is recommended.

H. S. V. K.

Testing chromium plate for resistance to abrasion. HARRY C. WOLFE, *Metals and Alloys* 2, 60-1(1931).

A. J. MONACK

Materials of construction in aircraft engines. R. R. MOORE, *Proc. Am. Soc. Testing Materials* 1930 (preprint) No. 33, 46-58.—A discussion of the various steels, bronzes, brasses and light alloys most commonly used in aircraft engine construction. Phys. properties and chem. analyses of the alloys are tabulated and the applications of each alloy and the reasons for its specific uses explained in detail.

W. H. BOWTOW

Ferrous metals used in airplane construction. J. B. JOHNSON, *Proc. Am. Soc. Testing Materials* 1930 (preprint) No. 33, 3-9.—Cast Fe and cast steel have but a limited application in aircraft construction. Standards on the chem. compns. as adopted by the Society of Automotive Engineers, in some cases with slight modifications, are applied in the airplane industry in the cases of rods, bars and billets. Rolled sheets and strips, tubes and wire are used in the industry.

W. H. BOWTOW

Some notes on teeming speeds of ingots. A. JACKSON, *Iron Steel Ind. and British Foundryman* 4, 167-70(1931).—A summary of the effects of teeming rates on 6-ton ingots. Rapid teeming rates tend to cause cracking.

C. H. LORIE

Multi-component systems involving iron. III. The system: iron-phosphorus-silicon. W. HUMMELTSCHE AND F. SAUERWALD, *Z. anorg. allgem. Chem.* 194, 113-38 (1930), cf. C. A. 24, 1049.—The system contg. an excess of Fe was exam'd chemically, thermally and metallographically. $\text{Fe}_2\text{P}-\text{FeSi}$ is quasi-binary. $\text{Fe}-\text{Fe}_2\text{Si}-\text{FeSi}$ is an independent 3-component system. $\text{Fe}_2\text{P}-\text{FeSi}$ is not quasi binary. A ternary eutectic at 1018° consisted of 6.35% Si, 7.45% P, 86.20% Fe. At 1110° there is an equl. of the form $\text{Liquidus } L + \text{Fe}_2\text{P} = \text{Fe}_2\text{P} + \text{FeSi}$. The liquidus is 13.55% Si, 6.00% P, 80.45% Fe. Eleven diagrams are given and discussed, also a no. of photomicrographs.

B. A. SOULA

The expansion of iron. W. SCHNECK, *Gieseler-Ztg.* 27, 1-3(1930).—Motor cylinders are cooled by water or air, but the pistons come in direct contact with hot gases and expand. On subsequent cooling the Fe does not contract to its original vol., thereby it suffers a loss of strength. The modern light construction demands stringent specifications. High Si causes pptn. of the C as graphite, with accompanying expansion. Mn, and especially Cr, hinder the pptn. In order to regulate the compn. accurately, the Fe should be melted in an elec. furnace.

CURTIS L. WILSON

The coefficient of expansion of cast iron. F. ROLL, *Gieseler-Ztg.* 27, 4-7(1930).—A review of the effects of temp. and alloying elements upon the coeff. of thermal expansion of cast Fe.

CURTIS L. WILSON

Shrinkage in cast iron in relation to the production of some special castings. R. T. ROLFE AND J. LAING, *Iron Steel Ind. Brit. Foundryman* 3, 263-7, 284(1930).

E. J. C.

Factors influencing the hardness of cast iron. AL. L. NORBURY. *Metallbörse* 20, 1939-40(1930).—The influence of various factors upon the pptn. of fine graphite particles in the molten Fe is discussed. The addn. of graphitizing substances probably causes the pptn. of a large no. of finely divided graphite particles which are dissolved easier and quicker than the big flakes. Gray castings made with the addn. of such graphitizing substances show better strength values. Fe, Si, Ca, S or Ni addns. have this effect. Mn and S by themselves favor the tendency toward hardening. Since they are always present at the same time, however, they neutralize each other through the formation of MnS. In order to neutralize the effect of S, an excess of 0.3% Mn over the quantity theoretically required for the formation of MnS is necessary. P seems to diminish the tendency toward hardening of thin pieces cast in sand molds. The tendency toward hardening is diminished by Al, Ni, Cu and Co, while it is increased by V, Cr, Se, Mo, Sn, Sh and W. LEOPOLD PESSEL

Effects on cast iron of prolonged heating at 800-1100°F. R. S. MACPHERLAN AND REXFORD H. KAUGER. *Trans. & Bull. Am. Foundrymen's Assoc.* 2, 828-64(1931).—See C. A. 25, 60. C. H. LORIO

Chromium in cast-iron mixtures. T. F. JENNINGS. *Trans. & Bull. Am. Foundrymen's Assoc.* 2, 801-9(1931).—See C. A. 24, 4743. C. H. LORIO

Effect of nickel and silicon on the properties of whiteheart malleable cast iron (with special reference to thin-walled malleable). I. S. J. E. DANGERFIELD, P. JOHNSON AND E. R. TAYLOR. *Iron and Steel Inst. (London), Carnegie Scholarship Mem.* 19, 1-18(1930).—The effect of Ni on cupola melted white cast iron used for thin-walled malleable castings was investigated. The test bars, $\frac{3}{16}$ in. thick, contained C 3.4, Si 0.65, S 0.2, P 0.05 and Mn 0.03%. The content of Ni varied from 0 to 3.66%. Ni improves the fluidity of the molten alloy. The tensile strength increases to a max. of 22.7 tons/sq. in. at 2.5% Ni and falls to 6.4 tons/sq. in. at 3.66% Ni. Bend tests give the best results at 0.7% Ni. The machinability is not impaired by the presence of Ni. In the annealed bars Ni causes a diminution in the size of the C nodules, and up to 2.8% a retardation in the rate of C migration, thus retaining more combined C in the form of an intercryst. network. This is attributed to a widening of the temp. interval of the Ar₁ range due to the Ni content. Above 2.5% Ni there is a rapid decrease in the percentage of combined and total C left in the annealed bars. With 3.66% Ni only 0.085% C is left and a severely burned material results. H. S. V. K.

The influence of phosphorus, sulfur, nickel and chromium on the growth of cast iron. O. BAUER AND K. SIPP. *Die Giesserei* 17, 989-95(1930).—Investigations to det. the influence of P, S, Ni and Cr on the primary growth (by annealing tests), and the carbide decompn. of the graphite sepn. from the solid soln. (metallographically) of east Fe are described. The growth is promoted by Si, P and Ni and is retarded by Cr, Mn and S (in large amts.), while S in small quantities and C have no influence. Below the pearlite point P promotes the carbide decompn., and at high temps. the C sepn. from the solid soln. Cr and Mn tend to retard both these, while S in small quantities does not influence the carbide decompn. In large amts. S tends to act strongly against both this decompn. and the C sepn. Micrographs (18) are given. J. BALOZIAN

Effects of alloys in cast iron. R. C. GOOD. *Trans. & Bull. Am. Foundrymen's Assoc.* 1, No. 8, 15-28(1930).—A review is given of the effects of alloying constituents. The influence of C, Si, P, S, Mn, Cr, Ni, Co, V, Ti, Al, Cu, Mo, W and Zr is discussed. LEOPOLD PESSEL

Some notes on cast-iron plating tools. ROGER POWELL. *Iron Steel Ind. and British Foundryman* 4, 161-2(1931).—From service records the essential of a good plating tool appears to be a good chill $1\frac{1}{2}$ to 3 in. deep. C. H. LORIO

Microstructure of pig and cast iron. H. PINAL. *Giesserei-Ztg* 27, 436-47(1930).—A series of 133 photomicrographs showing the various structures of P, graphite, combined C and slag in pig Fe and cast Fe. CURTIS L. WILSON

The properties of pig irons from various sources. A. WAGNER. *Giesserei-Ztg* 27, 403-12(1930).—Chem. analysis is not an exclusive criterion of the value of pig Fe, as many varieties of the same compn. exhibit different properties. The behavior of 17 types of pig Fe, such as English hematite, foundry and Swedish charcoal Fe was investigated by melting and remelting several times in crucibles, under as const. melting conditions as possible, and examg. specimens for tensile and bending strength, hardness and microstructure. The first melting resulted in materials of low tensile strength, 8.5-12 kg., and low bending strength, 18-19 kg. Repeated remelting improved the phys. properties: after the 4th melting, the tensile strength increased 100% without any appreciable change in compn. Numerous photomicrographs are reproduced. CURTIS L. WILSON

The causes of blisters and pinholes in wet-process cast-iron enamel. WALTER KERSTAN. *Die Giesserei* 17, 965-72(1930)—Plant and lab expts were carried out to det. the causes of blisters and pinholes in cast-Fe enamels produced by the wet process. These defects are shown to be due to the simultaneous action of various causes, principal among which are (1) incorrect compn of the frnt, (2) phys defects in the cast Fe, (3) insufficient cleaning of the casting surface, and (4) various changes in the Fe structure during the baking of the enamel. Cast Fe in which the graphite is finely divided is the most stable during baking and shows the least change in structure during the enameling process.

J. BALOZIAN

Etching figures appearing in iron and steel. V.-N. SVETNIKOV. *Rev metal* 27, 512(1930)

C. H. LORIO

The study of special steels—constitutional diagram of the 18 per cent Cr alloys with variable composition of nickel and carbon. V. N. KRIVONOK. *Mining Met Investigations U S Bur Mines, Carnegie Inst Tech, Mining Met Advisory Boards, 4th Open Meeting of the Met Board, Oct. 17, 1930, 4 pp*—An outline is given of problems begun or to be carried out on 18% Cr-Fe alloys contg variable percentages of Ni and C.

C. H. LORIO

Mechanical properties of copper steels with special reference to the effect of heat treatment. P. NUNN. *Stahl u Eisen* 50, 678-80(1930)—The addn of more than 0.6% Cu to mild steel increases the yield point appreciably but tends to make the metal brittle, a tendency which is overcome by the addn of 0.4% Cr. This alloy has a much higher tensile strength and yield point at high temps than Si steel and is more resistant to corrosion. Quenching from above 800° and tempering at 450-550° increases the tensile strength of steels with more than 0.6% Cu without seriously reducing the ductility and resistance to shock. Tables and graphs are given showing the effect of various heat treatments on the mech properties of several Cu steels after varying mech treatment.

B. C. A.

Vibratory strength of nitrided steel. O. HENOSTENBERG AND R. MAILANDER. *Krupp Monatsh* 11, 252-4(1930)—A considerably higher vibratory strength is shown by the nitrided specimens as compared with the same steels not nitrided. This depends on the relative depth of nitrided layer and core as well as upon the type of steel. The first breakdown always occurs at a point on the border of the nitrided layer, although this layer shows very little plastic deformation. Small surface defects are without effect upon the vibratory strength of nitrided samples.

H. STORATZ

Practical importance of the influence of different cooling conditions on the structure of steel ingots. P. LEITNER. *Stahl u Eisen* 50, 1081-6(1930)—The effect of varying wall thicknesses of the mold on the structure of plain C and Ni-Cr steels has been examined. The results show that a more fine grained and homogeneous macrostructure is obtained by retarded cooling and hence thin walled molds are to be preferred. Even better structures are produced by retarding the radiation of heat from the mold by covering it with a heat insulating material or by using a double-walled mold with an annular air space between the walls.

B. C. A.

Drawing of rapidly quenched steel. ANDRÉ MICHEL AND PIERRE BENAÏET. *Compt rend* 192, 163-6(1931)—A study is made of the influence of drawing time and temp on the transformation of austenite to martensite in a high speed steel contg C 0.72, Si 0.33, Mn 0.26, Cr 5.20, W 18.33, V 1.02 and Co 11.09%. Co in steel increases the effect of secondary tempering.

C. H. LORIO

Drawing of special rapidly quenched steels. ANDRÉ MICHEL AND PIERRE BENAÏET. *Rev metal* 27, 501-8(1930)—Dilatation characteristics of 16 alloy steels were detd to study the $\gamma \rightarrow \alpha$ transformation.

C. H. LORIO

The annealing of steel. P. CHEVENARD. *Metaliborse* 20, 1828(1930)—The hardening of steel was investigated by means of photomicrographic, magnetic and dilatometric methods. Dilatometric curves were detd for the thermic cycles of growth with gradually increasing temps and also for the phenomena of isothermic growth under exact detn of the change of length as a function of the time. A special dilatometer with movable photographic plates was used. Constant temp was insured by special construction. The investigations were made on austenite, the only constituent that can be obtained in a pure state. It was investigated in the state of complete superhardening. A mixt of austenite martensite was investigated in a partially superhardened condition and the greatest hardness obtainable was detd in a quenched steel contg the max amt of martensite. The phenomena can be explained by the equal diagram but they are complicated by the simultaneous occurrence of the various reactions. C could eliminate the influence of the simultaneous reactions which had caused the variations of results before. Austenite can lead to the formation of cementite.

a reaction which causes contraction, or a mixt. of α -Fe and cementite may be formed under expansion. The formation of martensite, which follows that of austenite, is connected with a contraction that grows with the temp. In quenched steels there is always a simultaneity of these 2 phenomena, causing in certain sections of the expansion curve either an arrest in the contraction, a rise or a slowing down. A direct transition of austenite into martensite during heating could never be observed. This secondary reaction in hardening always occurs in the second period of cooling. This simultaneity of the phenomena in martensite and austenite can also be observed by means of the elec. cond. LEOPOLD PESSEL.

Solution of scale in pickling steel. E. L. CHAPPELL AND PAUL C. ELY. *Ind. Eng. Chem.* 22, 1200-1 (1930).—In pickling welded pipe 95% or more of the scale goes into soln., the undissolved part consisting mainly of Fe silicates. The rate of scale soln. is about equal to the rate of metal soln. in the absence of an inhibitor. The rate of scale soln., is about the same with or without an inhibitor. Scale on pipe is more quickly and completely dissolved than scale in acid without metal. L. P.

Pickling solution for stainless and nickel steel. W. E. WARNER. *Machinery* 36, 626-7 (1930). *Metals and Alloys* 1, 737.—Proportions given are: 9 parts water, 1 part H_2SO_4 , 1 part HCl and 1 part HNO_3 . G. G.

Economic pickling of thin metallic sheets. E. DWORZAK. *Przemysl Chem.* 14, 361-73 (1930).—The rate of soln. of thin sheets of Fe and Zn in 1-25% H_2SO_4 at 30-50° was studied with and without the addn. of certain inhibitors. This rate for Fe was followed by measuring the vol. of evolved H. In H_2SO_4 alone the rate of soln. of Fe increases rapidly with temp. At 30° this rate is markedly greater in 35% than in 5% acid, but is nearly the same for all concns. at 50°, when 1 cc. H is evolved per 1 cm² of the surface per 20 min. The addn. of 1-15% naphthalenesulfonic acid (I) or "Passiv" to 15% H_2SO_4 reduces the rate of evolution of H markedly and causes the rates for various concns. of acid to remain different even at the higher temps. The effect of the addn. of 0.1-2% "Vogel's compd." or "compd. K" (author's prepn.) is to repress the rate of H evolution still further, although the former ceases to function at 50° after 30-40 min., while the latter continues to maintain its inhibiting property. The advantage of speeding up the pickling of Fe by increasing the temp. or concn. of the acid is only apparent, because the increased rate of H evolution is not only a sign of an increased rate of consumption of Fe and acid but is also harmful to the metal. The pickling process is speeded up appreciably by the addn. of inhibitors. Their concn. seems to have little effect on the rate of H evolution at the higher temps. H_2SO_4 alone removes the oxide scale on Fe sheet in small chunks. In the presence of I these chunks are larger. In the presence of "Vogel's compd." or "compd. K" there is no H₂ observed at all until after several min. the whole surface of the oxides suddenly peels off. A study of the rate of soln. of Fe oxide scale in H_2SO_4 in the presence of the above inhibitors leads D. to conclude that evidently the negative catalytic effect of the inhibitors pertains to the soln. of metals in acids, and of those reactions which are accompanied by evolution of H, and not to the metal oxides. Tests show that the inhibitors increase the rate of soln. of the scale, although in com. practice only a negligible amt. of scale is dissolved. In pickling, the scale is forced away from the metal by H which is produced by the action of the acid on the metal. A study of the efficiency of the utilization of acid showed that 40 kg. 60° Bé H_2SO_4 is needed for pickling 1 T iron sheet, 1 sq. m. surface and 1 mm. thick in 15% acid, whereupon 2.3% metallic Fe is dissolved. The addn. of 3% "Passiv" reduces this to 1.06%. The corresponding figure for "Vogel's compd." is 0.78%, and for "compd. K" 0.53%. A. C. ZACHLIN.

Pickling of pipe using commercial inhibitors. E. L. CHAPPELL AND PAUL C. ELY. *Ind. Eng. Chem.* 22, 1201-3 (1930).—The use, lab. control and practical evaluation of inhibitors are discussed. LEOPOLD PESSEL.

General principles underlying the use of alloy steel. R. WHITFIELD. *Iron and Steel Ind. and British Foundryman* 3, 347-50 (1930).—A review. LEOPOLD PESSEL.

The constitution of alloys. The application of physical methods. ALFRED SCHULZE. *Giesserei Ztg.* 27, 86-97 (1930).—Of all the properties used to study the constitution of alloys, the elec. cond. is the easiest to comprehend and the most sensitive toward slight changes in constitution. The change in elec. cond. for binary systems which show complete insoly. in the solid state, when plotted against the vol. concn., is not quite linear; it is not the arithmetic mean, but the geometric mean, of the elec. conds. (and elec. resistivities). For complete miscibility in the solid state, the curve of elec. cond. decreases rapidly and has in the middle a flat min. The resistance resulting from solid solns. is independent of the temp. When a system shows partial soly. in the solid state, the elec. cond. curve decreases rapidly within the limits of soly. and then changes almost

along a straight line. When a chem compd is formed, a distinct break in the elec. cond. curve is noted at the concn of the compd. The relation between concn., temp., and elec. cond. is shown on space models. CURTIS L. WILSON

The thermal properties of various alloy groups. ALFRED SCHULZE. *Gieseler-Ztg.* 27, 177-84(1930), cf. preceding abstr.—Next to the elec. cond., the thermal expansion of alloys throws the most light upon their constitution. The tube method of dilatometric analysis is both sensitive and exact. For systems completely insol. in the solid state, the coeff. of thermal expansion is a linear function of the vol. concn. For complete soly in the solid state, the curve of the coeff. of expansion is convex toward the concn. axis. This deviation from a straight line is explained as due to the fact that an atom of the one component displacing an atom of the other disturbs the space lattice so that the av. distance between the atoms is not proportional to the concn. For partial soly systems the coeff. curve is similar to that of solid solns. within the soly limits, outside of which it is a straight line. The thermal cond. is similar to the elec. cond. with respect to its relation to the constitution of alloys. Magnetic and x ray analyses are also discussed. CURTIS L. WILSON

Structural and engineering light alloys for aircraft. R. L. TEMPLEY, F. V. HARTMAN AND E. C. HARTMANN. *Proc. Am. Soc. Testing Materials* 1930 (preprint) No. 33, 10-6.—The wrought and cast alloys of Al and Mg best suited to aircraft work and their respective advantages are discussed. The designation of the various alloys available is indicated and the forms in which they are produced commercially given. Data on nominal compn., heat treatment and mech. properties are tabulated. W. H. B.

Alloys for drilling purposes. N. PRADYUK AND V. NEKREEV. *Azerbaidzhanskoe Neftyanoe Khozyaistvo* 1930, No. 12, 94-112, cf. C. A. 24, 2406.—Analyses and descriptions are given of various alloys which were tested for hardness and resistance to wear. Expts. indicate that alloys contg. W can be successfully applied to drilling. V. K.

Alloyed castings. E. KOTHNY. *Gieseler-Ztg.* 27, 291-300, 323-7(1930).—A review of the effects of Ni, Cr, Ni-Cr, Ti, Al, B, Ce, Co, Cu, W, Mo, V and Zr upon the constitution and the phys. and chem. properties of gray, chilled and malleable Fe castings. A bibliography is appended. CURTIS L. WILSON

Processes involved in age-hardening. G. TAMMANY. *Z. Metallkunde* 22, 305-8(1930).—The theory of age hardening is discussed generally. It is concluded that the increase in hardness and strength observed in age hardening alloys cannot be explained by the blocking of slip-planes by ppt. alone, but that the chief effect must be a pre-pptn. assocn. of the pptg. mols. in the supersatd. solid soln. The importance of the form of the ppt. (Widmanstätten figure) and of the process of coagulation is discussed. R. F. MEHL

Age-hardening of alloys. W. GUERTLER. *Z. Metallkunde* 22, 78-84(1930).—A review, chiefly of the theoretical aspects. ROBERT F. MEHL

The age-hardening of Constructal. C. WASSERMANN. *Z. Metallkunde* 22, 160-2(1930).—The changes in lattice parameter during aging of the alloy Constructal (Al with 5.88% Zn, 1.40% Mg, 1.35% Mn, 0.28% Fe, 0.20% Si) was studied. It was found that the lattice const., a_0 (the side of the unit face centered cube of the Al solid soln.) decreased by 0.0015 Å. U. upon room temp. aging for one day or artificial aging at 80° for 1/2 hr. Annealed wire showed the presence on the x ray spectrogram of $MgZn_2$. The strongest line of $MgZn_2$ may be found on the spectrogram after only 1/2 hour's aging at 80°. It was not possible to prove the presence of $MgZn_2$ ppt. in alloys aged only at room temp. ROBERT F. MEHL

Brinell hardness, elastic limit, and ultimate tensile strength of age-hardening aluminum alloys. P. MELCHROX. *Z. Metallkunde* 22, 175-6(1930).—The equations suggested by Bohner (C. A. 24, 1007) to express empirical relationship between these quantities are discussed and simpler and more accurate equations proposed. ROBERT F. MEHL

The hardening problem in copper-silver, beryllium-copper and zinc-copper alloys. M. HAAS AND D. UNO. *Z. Metallkunde* 22, 154-8(1930).—Cu-Ag.—Earlier work on the solid soly of Cu in Ag is summarized. A new solid soly curve is established by measurements of thermal expansion (dilatometric) and of elec. resistance and by microscopic studies on quenched and annealed alloys. The maximums upon the dilatometric curve and the elec. resistance curve correspond to the max. upon the hardness curve. Be-Cu.—Dilatometric curves on the Cu rich solid soln. are given. Hardness measurements are in good agreement with the dilatometric. Zn-Cu.—A few dilatometric curves are given for alloys cooled from the β field. Throughout this work an attempt is made to trace the course of pptn. by means of changes in volume and to correlate these changes

and therefore the various stages in the pptn process with the changes in hardness observed
ROBERT F. MEHL

The alloys of aluminum. L. POTIN. *Rev. gen. sci.* 41, 474-7 (1930).—A review
LEOPOLD PESSERL

Cast aluminum bronzes. III. Shrinkage phenomena. W. CLAUS AND F. GOEDERITZ. *Die Gießerei* 17, 1017-23 (1930), cl. C. A. 24, 4197.—In studying the microstructure of cast Al bronzes, castings of 4.5% and 10% Al, and a Si Al bronze are made in dry and moist sand and in chill molds pouring at 1250°. Shrinkage pipe is present in the core of the 7 and 10% Al bronzes when cast in either the dry or the moist molds. The 4.5% Al bronze shows small defects in the bottom of the casting, but no shrinkage pipe in the head. This bronze when cast in a chill mold gives a dense casting. Shrinkage pipes are shown by the 4.5, 7 and 10% bronzes when cast in chill molds. The Si Al bronze shows slag defects in the foot when cast in the sand molds, and a central porosity when cast in the chill mold. Lower pouring temps. (1050°) do not affect the shrinkage pipe in the 10% Al bronze when cast in the dry sand mold.
J. BALOZIAN

Ternary alloys of aluminum, silicon and copper. G. G. URAZOV, S. A. POGODIN AND G. M. ZAMORUEV. *Mineraln. Suv. i Tsvetn. Met.* 4, 160-65 (1929).—A study, by thermal analysis, of the Al triangle up to 21% Si and 40% Cu is given. The alloys in this range consist of mech. mixts. of Si-CuAl₃ and a solid soln. of Si and Cu in Al. The soly. of Cu is increased by Si in the interval 450-350°. The hardness of the ternary solid solns. was detd. Aging of the alloys at the ordinary temp. diminishes with increasing content of Si. The technical application of the results is considered.
B. C. A.

Bismuth alloys. J. G. THOMPSON. *Metals and Alloys* 2, 92-4 (1931).—See C. A. 25, 670.
A. J. MONACK

Crystal structure of the alloys of iron and manganese. JAMES B. FRIEDL AND MAXWELL GENSAMER. *Mining Met. Investigations U. S. Bur. Mines, Carnegie Inst. Tech., Mining Met. Advisory Board, 4th Open Meeting of the Met. Board*, Oct. 17, 1930, 6 pp.—X-ray analyses show α -Fe is present in Fe-Mn alloys up to at least 16% Mn. γ -Fe makes its appearance at room temp. in the forged alloys with 13% Mn. Another phase, the hexagonal close-packed arrangement, was found in forged alloys contg. 10 to 26% Mn.
C. H. LORIE

Constitutional diagram of iron-manganese alloys. V. N. KRIVOBOK. *Mining Met. Investigations U. S. Bur. Mines, Carnegie Inst. Tech., Mining Met. Advisory Board, 4th Open Meeting of the Met. Board*, Oct. 17, 1930, 3 pp.—A tentative constitutional diagram was constructed from data supplied by thermal analysis, metallographic studies and x-ray patterns. It is believed there are at least 6 different phases: α -Fe, γ -Fe, ϵ -Fe, α -Mn, β -Mn and γ -Mn.
C. H. LORIE

Heat treatment and metallography of the alloys of iron and manganese. V. N. KRIVOBOK AND CYRIL WELLS. *Mining Met. Investigations U. S. Bur. Mines, Carnegie Inst. Tech., Mining Met. Advisory Board, 4th Open Meeting of the Met. Board*, Oct. 17, 1930, 4 pp.—Dividing the Fe-Mn alloys into 6 groups, K. and W. discuss the microstructure of each group individually but briefly.
C. H. LORIE

Thermal analysis of iron-manganese alloys. F. M. WALTERS, JR., AND CYRIL WELLS. *Mining Met. Investigations U. S. Bur. Mines, Carnegie Inst. Tech., Mining Met. Advisory Board, 4th Open Meeting of the Met. Board*, Oct. 17, 1930, 6 pp.—Thermal analyses were made in a gradient furnace on Fe-Mn alloys contg. 17, 44, 72, 102, 131 and 16% Mn with an av. of 0.02% C. A sudden change in thermal disturbance was observed in alloys contg. 13 and 16% Mn.
C. H. LORIE

Ternary alloys of lead. S. J. NIGHTINGALE. *Brit. Non-Ferrous Metals Research Assoc. Develop. Rept.* D3, 12 pp. (1929).—The properties of Pb alloys contg. (a) 0.25% Cd and 0.5% Sb and (b) 0.25% Cd and 1.5% Sn are compared with those of pure Pb as regards their utility for the manu. of pipe. The tensile strength of the alloys is 1.62 tons/sq. in compared with 0.9 ton/sq. in for pure Pb, the ductility 52-55% against 80% and the fatigue limit 0.74-0.57 ton/sq. in against 0.18. The alloys are on the whole more resistant to corrosion than is Pb and withstand a considerably higher bursting pressure, so that a lighter pipe can be used for the same service than is the case with Pb. The working and soldering properties of the alloys are equally as good as those of lead and, taking into account their superior properties, are relatively less costly than Pb; they do not undergo intercrystalline corrosion or fatigue.
B. C. A.

Silicon-zinc-copper alloy. E. VADERS. *Metal. Ind.* (N. Y.) 29, 108-10 (1931).—
See C. A. 25, 63.
G. G.

Useful titanium alloys. W. KROLL. *Metallwirtschaft* 9, 1043-5 (1930).—Fe-Ti

alloys can be rolled and may be used as alloy steels. With high Ti content they are susceptible to aging. Their grain structure is very coarse. Ni reduces grain size and increases the age hardening effects. The alloys are compared to those of Fe-Bc.

C. H. LORIG

Directional properties in cold-rolled and annealed copper. ARTHUR PHILLIPS AND F. S. BURN. *Am Inst Mining and Met Eng Tech Pub No 413*, 18 pp (1931).—The paper deals with the directional differences in the tensile properties of 2 kinds of sheet Cu: tough pitch Cu contg 0.0295% Ag and electrolytic Cu deoxidized with P (residual P 0.0089%). The first series of tests was on the cold rolled Cu, reduced 10–90%. Tensile tests were made on strip cut (1) in the direction of rolling, (2) at 90 degrees and (3) at 45 degrees to the direction of rolling. In the second series the 3 kinds of strips cut from the rolled sheets were annealed over a temp. range extending from 300° to 800°. The results seem to indicate that rather pronounced directional differences, particularly in elongation values, are produced by the combined effects of heavy reductions and high temps. of anneal.

C. L. MANTLE

Duplexing. K. R. BRINKA. *Iron and Steel Ind and Brit Foundryman* 4, 194, 193 (1931).—Processes are described.

C. H. LORIG

Parkerizing. E. LIEBKNECHT. *Z. anorg Chem* 43, 769–71 (1930).—The reactions underlying this process and the qualities of the coatings obtained are discussed.

LEOPOLD PESSER

An inquiry into lead poisoning during coating with lead on metal. E. AGASSE-LARONT. *Med Travail* 1, 277–9 (1930). *Bull Hyg* 5, 923–4.—Metals are coated by plunging the red hot metal into the metal powder or by sprinkling the powder on the hot metal and returning the metal to a hot furnace where it melts and spreads. The powder may contain 66% of Pb_3O_4 . The air is so laden with the powder that 6.6 g. (contg about 4 g. of Pb) were collected on each sq m. of surface per hr. Basophile granulated red cells were found in the blood of 90% and 80% showed the blue line. They suffered from colic. Symptoms of plumbism were present after a few days' employment.

GEORGE R. GREENBANK

Risks in new methods of spraying. W. ROSENTHAL, D. ROSENTHAL AND ERIKA. *Deut Med Wochschr* 56, 1444–67, 1487–9 (1930). *Bull Hyg* 5, 920.—The authors describe "metallization" in which molten metal is blown onto a surface by compressed air. Cellulose spraying is described in which turpentine, C_6H_6 , and its much less dangerous homologs $PhMe$ and C_6H_5Me , are used. Particular stress is laid on the dangers of $MeOH$ as likely to produce optic nerve neuritis. Among the poisonous coloring matters discussed Pb and Si are the most dangerous. The value of periodic medical exams is stressed, and the symptoms which if present should result in suspension of the worker are indicated. Symptoms of poisoning by C_6H_6 , Pb and Si are described.

GEORGE R. GREENBANK

Protecting iron from rust. DIECKMANN. *Farbe u. Lack* 1931, 87–8.—An address.

G. G. SWARD

Corrosion-resisting cast iron. W. ACKERMANN. *Chimieri-Ztg* 27, 203–6 (1930).—As corrosion is regarded as an electrochem. process the homogeneous structure of solid solids is most favorable to resistance. Graphite accelerates corrosion. Si causes the pptn. of graphite, but a cast Fe with 0.84% Si corrodes more rapidly than one with 2.25% Si. The resistance of the higher Si casting is attributed not to the graphite but to the Si which is in solid soln. P up to 0.5% increases resistance, but more than 0.07% S is detrimental. From 14 to 22% Si creates an acid resisting material, which is not recommended for alk. solns. Ni, Cu and Cr reduce the rate of corrosion under certain conditions. The addn. of Monel metal to cast Fe produces an austenitic structure and a resistance to corrosion almost equal to that of bronze.

C. L. W.

Corrosion-prevention methods as applied in aircraft construction. H. S. RAWDON. *Proc Am Soc Testing Materials* 1930 (preprint) No 33, 30–6.—A résumé of present com. practice on corrosion prevention. Steel structural parts are plated either with Cd or Zn followed by an iron-oxide primer and a lacquer or enamel. The inside of closed tubular members is coated with the primer. In Navy practice duralumin is treated by the anodic-oxidation process prior to the application of a priming coat, followed by the finishing coat. Army practice is to have a careful preliminary chem. surface cleaning followed by a priming coat and finishing coats. Special precautions against "crevice" corrosion involve the use of rubberized balloon tape, bituminastic paint or similar product between metal parts and wood, leather, etc., and the use of "dope proof" paint or Al foil on top of the protective coating already applied. Duralumin and steel parts before assembly by riveting or bolting are coated with primer or Navy gray enamel on the contact surfaces. Bare Al surfaces are cleaned by the use of a

mild etching soln., such as a dil. soln. of H_2PO_4 . Another soln. is a caustic "dip" followed by neutralization with weak acid. Mild abrasives may be used, even sand blasting for heavier parts, each followed by chem. cleaning. W. H. ROYSTON.

Corrosion exfoliations on aluminum sheet. LEONARD PRINCE. *Metal Ind.* (N. Y.) 20, 72-3 (1931).—The phenomenon of corrosion exfoliations, which is quite different from the commonly observed forms of corrosion of Al, can be observed on Al sheet exposed to continuous salt spray. This phenomenon is explained on the basis of the O'Brien cell theory, under the assumption that min. crevices, formed as a consequence of the rolling process, are made the anodic areas because of the exclusion of O. The formation of a large vol. of Al hydroxide takes place in these sections causing the blister-like rising of the adjoining metal. It was shown by expt. that Al sheet, sand blasted and sprayed with molten Al by the metal spraying process, will not show this type of corrosion. LEONARD PRINCE.

The corrosion of copper and high-copper alloys by salt solutions with special consideration of conditions existing in the potash industry. A. RAUCH AND H. KOLN. *Korrosion Metallschutz* 6, 151-6, 174-8, 193 200 (1930), cf. C. I. 24, 4750. Corrosion time curves obtained by the corrosion of Cu in 1.35 N KCl and 1.35 N $MgCl_2$ solns., while of a linear nature, show a kink after approx. 1 hr., coinciding with the time when a deposit of basic Cu chloride could be observed. Solns. of Na_2SO_4 and $MgSO_4$ were much less active toward Cu than the chlorides. The curves obtained show considerable spreads, making it difficult to draw a definite conclusion. A comparison between electrolytic Cu and 4% Al bronze showed that the linear relation between time and corrosion is found with Cu alloys, also. While the corrosion was greater with chlorides than with sulfates, there was also found a distinct relation between the nature of the anion and corrosion. In different salts having the same anion, the corrosion was the same at equal concn. of this anion. With increasing velocity of the liquids a slow increase of corrosion could be ascertained. The curve apparently approached a max. An app. is described in which the corrosion of Cu was accomplished in the absence of O. As was to be expected, the corrosion in solns. free of O was found to be only a fraction of that observed in the presence of air. In the absence of O the corrosion is practically not influenced by changes in concn., temp. and time of exposure. It is also equally intensive with chlorides and sulfates. In the absence of air, pure Cu and Cu alloys show the same behavior. Here too, no essential influence of concn., time and temp. could be found. Technical salt solns. were found to give practically the same result. The small difference in the corrosion values found in the presence and absence of air can be explained by the low O content of the solns. and by the tenacious retention of O in these concd. solns. For the detn. of O in salt solns. Winkler's Mn hydroxide method was modified for solns. contg. Mg salts by using NH_3 contg. NH_4I and NH_4Cl in place of NaOH. This method gave good results. The concn. of O in salt solns. was found to decrease with increasing concn. The influence of the temp. upon the O concn. was detd. by modifying the app. The solv. of O decreases with increasing temp., the relation being expressed by a linear curve, apparently up to the h. p. of the solns. No quant. relation could be found between O solv. and corrosion. The corrosion reaches a max. at 20% KCl. From there on influences favoring and counteracting corrosion are apparently active at the same time. Leaching solns. attack Cu much less than KCl solns. of moderate concn.; this is explained by the lower solv. of O in the former. The temp. coeff. of corrosion in leaching solns. is also smaller. An experiment made to ascertain the attack by carried-over solid salts indicated the negligible degree of this factor. Solid $MgCl_2$, however, can, if in contact with Cu at increased temp., induce corrosion by splitting off HCl. The extent of corrosion of various Cu-bearing alloys was investigated. Al bronze was generally found superior to smelters Cu, electrolytic Cu, tin bronze and brass. While the latter showed greater resistance in one expt. with flowing KCl solns., the danger of severe local dezincification exists and suggests Al bronze as the more advisable material. Smelters Cu was found to be least resistant; electrolytic Cu came next, except in higher KCl concns., where Sn bronze showed greater attack. In an investigation of etched and copper-plated Cu sheets the influence of the surface condition upon corrosion was found to be negligible. Cold-worked Cu and Cu alloys were found to be more attacked than the soft annealed material. Potential measurements were made on electrolytic Cu, Al bronze and Sn bronze along with detns. of the corrodibility. No connection was found between the latter and the electro-chem. behavior. Electrolytic Cu, which showed the greatest corrosion, always had the noblest potential. In comparing the potentials of "soft" and "hard" electrolytic Cu with the corrodibility, it was found that the hardened metal, although of a lower potential, was less attacked. A soft and a hard Al bronze showed

the same amount of corrosion and yet a distinct difference in the potential. In comparing normal and cold worked metal sheets the latter always showed a more negative potential and also greater corrosion. By direct measurement of electrolytic Cu and Al bronze electrodes, the latter were always found to be anodic. The conclusion is that the potentials cannot be considered a general indication of the corrosion resistance of the material. Electrolytic Cu became more anodic with increasing concn. of KCl. Expts. were made to determine whether Toedt's methods of measuring corrodibility by residual currents could be applied in the case of electrolytic Cu and Al bronze. The combination Pt-Al bronze gives a current 2-3 times as strong as the combination Pt-electrolytic Cu, which stands in contradiction to the repeatedly detd. greater corrosion resistance of Al bronze. In other expts the quantity of residual current and the corrodibility were detd. at the same time. The latter was found to be 2-8 times as great as the values calcd. from the quantity of current. In other expts the quantity of current was found to be the smaller the greater the loss of weight due to corrosion. No definite regularity can as yet be found in the relation between residual current and corrosion. Corrosion expts. were made with a larger exptl. installation. An app. recirculating 150 l. of leaching soln. is described. Al bronze was found to have the greatest resistance against corrosion, being about 7 times better than electrolytic Cu. In the expts. on a smaller scale this proportion had been found to be less pronounced. Sn bronze of 1.5% Sn showed a corrodibility similar to Cu. LEOPOLD FESSEL.

Welding aluminum. FARN GROVE-PALMER. *Metal Ind.* (London) 33, 509-10, 533-4, 546 (1928) E. J. C.

Soldering and welding aluminum and its alloys. ERNST WERNER. *Gusserei-Ztg.* 27, 72-6 (1930).—Al and its alloys are difficult to solder because (1) the thin skin of thermally resistant and chemically inert oxide hinders a firm alloying of the metal and solder, (2) the high sp. heat and thermal cond. of Al cause in spots a "quenching" of the liquid solder and (3) the strong electropos. nature of the metal sets up reactions in the presence of liquids or moist air. Most modern Al solders consist principally of Sn with some Al and Zn. Many contain also Sb, Bi, Cd, Cu, Pb, Ni, Ag and other metals. Soldered joints are tested in steam, boiling salt soln. or boiling water for several days. The tensile strength is 5 kg./sq. mm., which is lower than that of Al. In order to meet them and meet specifications, autogenous welding is recommended. Al is one of the most weldable of metals, if the following characteristics are kept in mind: (1) low resistance against oxidation, (2) high coeff. of expansion, (3) high thermal cond., (4) low m. p., (5) low tensile strength at high temps., (6) tendency to warp. Modern fluxes consist of alk. chlorides, fluorides, bisulfates etc. A slightly reducing oxyacetylene flame gives good results. Welded joints should be guarded against too rapid cooling. CURTIS L. WILSON.

Reaction between metal receptacles and foodstuffs (BLEYER, SCHWABOLD) 12. Röntgenographic investigation of the system Cd-Mg (DEHLINGER) 2. Effect of internal stresses on the magnetic susceptibility of metals (HONDA, SHIMIZU) 2. The reactivity of metallurgical coke (DEFRAYE) 21. Present position of the metallurgical coke industry in Central Europe (BERTHELOT) 21. A device for determining work input to a laboratory ball mill (GROSS, ZIMMERLEY) 1. Magnetic susceptibility and chemical investigation (SPENCER) 2. Surface protection of the light metals (LEIS, KOLKE) 26. Blast-furnace-slag portland cement (TOMARCHIO) 20. A recording dust concentration meter and its application to the blast furnace (SIMON, *et al.*) 1. Equilibrium diagrams of the Al-Mn, Cu-Mn and Fe-Mn systems (ISHIWARA) 2. The equilibrium Fe-Fe₃C-O (PINCALT) 2. The heat of mixture in molten metals (KAWAKAMI) 2. Cr oxide [for metallurgical purposes] (Brit. pat. 336,671) 18. Washer for ores (Ger. pat. 516,145) 21. Clarifying plant for ore industry (Ger. pat. 516,454) 1. Alloy armouring for electric cables (Brit. pat. 336,408) 4.

Aluminum: Facts and Figures. London: British Aluminium Co. Gratis. Reviewed in *Foundry Trade J.* 44, 161 (1931).

FABIAN, V. Praktische Härtereifibel. Aus der Praxis des Härtens. Berlin: Verlag Chemie G. m. b. H. 92 pp. M. 6. Reviewed in *Chimie & Industrie* 24, 1031 (1930).

GIERDZIEJEWSKI, K. Kurs Odlewnictwa. [Foundry Course.] Tom. I. Warsaw: Księgarni Techniczne. Reviewed in *Foundry Trade J.* 44, 130 (1931).

KELLER, JOHN F. Lectures on Steel and Its Treatment. 2nd ed., enlarged. Cleveland: Am. Soc. for Steel Treating. 323 pp. \$3.50. Reviewed in *J. Western Soc. Eng.* 36, 60 (1931).

Metalli leggeri e loro applicazioni (*New journal*) Published monthly by Soc. an. Umberto Polacco & Cia., Via Borgospesso 19, Milan. Vol 1, Nos 1 and 2 are dated Jan.-Feb., 1931. Price, L. 120

Powdered and Granulated Aluminium. London British Aluminium Co. Gratis. **SOMMES, A.** *La Lorraine métallurgique.* Paris Berger-Levrault 249 pp. Reviewed in *Chimie & Industrie* 24, 1030(1930)

Spektrographische Ausrichtungen für die metallurgische Analyse. Anleitung zur Wahl geeigneter spektrographischer Apparate für den Gebrauch in chemisch technischen Laboratorien. London Adam Hilger, Ltd

Statistische Zusammenstellungen über Aluminium, Blei, Kupfer, Nickel, Quecksilber, Silber, Zink und Zinn. 31st year Frankfurt C Adelman A G. 91 pp

Froth-flotation concentration of ores. PRACY C WRIGHT (to The Minerals Separation North American Corp.) Can 309,071, Mar 3, 1931 Ore in the form of a non-acid pulp is subjected to a froth flotation operation in the presence of dioxanthogen

Mat (including a perforated sheet of rubber) for distributing air through liquids in ore flotation or other processes. WALTER H CURRY and HENRY F DIETZ (to Hewitt-Gutta Percha Rubber Corp.) U S 1,792,285-6, Feb 10 Structural features

Grading ores by weight for air blast hearths. REMBRANDT FRALK, WM S DAVIES and WM S WALLACE Ger 513,787, Aug 18, 1923 Details of the app are given

Values from ores by leaching and electrodeposition R SANZ CARRERAS, Brit. 336,584, July 10, 1929 Ores contg Cu, Ni and Co are leached with a dil H_2SO_4 soln contg NaCl (suitably about 5%) Any insol portion of the ore, which may contain arsenides and sulfocyanides, is oxidized in a furnace and again leached, and the solns obtained may be further treated together or separately The liquid, about 13-15°Bé., is electrolyzed for Cu recovery (various details of app and procedure being given) and the electrolyte is then evapd to 28°Bé and cooled to ppt all earth metal and Mg salts It is further concd to 38°Bé and cooled to crystallize Co sulfate, which may then be dissolved and subjected to a special electrolytic treatment to deposit Co in scales Ni is recovered from the remaining electrolyte

Apparatus for separating ore constituents by air currents. ALANSON O TAYLOR, U. S. 1,792,404, Feb 10 Various structural details are described

Treating ores of metals. MEYER MINERAL SEPARATION CO Brit. 336,770, Sept. 30, 1929 Metalliferous material is heated in the presence of a so-called "reagent metal" (defined as a metal capable of existing in 2 or more states of oxidation, such as Fe, Cu, Ni or Mn) to an elevated temp somewhat below the decompn temp of the sulfate of the "reagent metal" and treated alternately with an oxidizing gas such as air and with an atm contg SO_2 and air The process is applicable to oxidized or sulfide ores contg Au, Ag, Pb, Ni, Mn, Cu, Co, Bi, Zn, Cd, Fe and similar heavy metals (which may themselves contain sufficient of the "reagent metal" or may be mixed with pyrites to supply the latter). If the ore is not amenable to sulfating, a halide such as NaCl may be gradually added to form chlorides Various details and examples are given.

Treating sulfidic ores. E J KOHLMEYER, Brit. 336,670, July 19, 1929 Sulfidic ores of metals such as Zn, Sn, Pb, Bi, Cd, As and Sb are blast smelted in a rotary furnace of such (described) construction that the flame is compelled to enter and leave by the same side, and air is blown on or into the melt until oxidation of the Fe present commences. In treating ore contg pyrites and ZnS , the oxidation to ZnO supplies sufficient heat that no further external heating is required A small proportion of lime may be added to form Ca ferrite with any Fe oxide produced and prevent stiffening of the melt, and a small proportion of coal may be added to prevent the Ca ferrite from taking up ZnO When Fe and Cu are present, a slag contg Cu Ca ferrite is formed; the Cu may be sepd as a mat from the slag by treatment with added S-bearing material such as pyrites.

Treating chromium ores. I. G FARBEYND, A.-G. Brit. 336,970, July 22, 1929. Cr ores, such as those which are to be worked up with an alkali carbonate in the presence of lime, dolomite or bauxite, are preliminarily heated in an oxidizing atm. to above 800°, with or without adding a small proportion of an oxidizing agent such as $NaNO_3$ or KNO_3 .

Treating titanium ores. MONTAN-UND INDUSTRIALWERKE FORM. JOH D. STARCK. (Willy Glaser, inventor). Austrian 120,673, Oct 15, 1929 Ilmenite, rutile and like ores are heated in an autoclave under pressure with an excess of H_2SO_4 . The reaction product is allowed to settle, and the liquid, comprising the excess of acid with some dissolved Ti and Fe, is sepd and used in treating a fresh batch of ore The solid resi-

due, comprising metatitanic acid and ferro-ferric sulfate, is sepd into its components by extn with water or mother liquor from a previous operation

Treatment of zinc-bearing material. WILLIAM C. HOORY (to The New Jersey Zinc Co.) Can 308,892, Feb. 24, 1931 ZnS concentrates contg insol compds of Mg and Ca are roasted to convert the Mg and Ca into sulfates and the ZnS into ZnO, and the product is leached to remove the sulfates of Mg and Ca

Treatment of zinc-bearing material. WILLIAM C. HOORY (to The New Jersey Zinc Co.) Can 308,893, Feb. 24, 1931. A pulp of Zn concentrates contaminated with insol compds of Mg and Ca in objectionable amts is treated with H_2SO_4 in amt. sufficient to convert the compds of Mg and Ca to Mg and Ca sulfates while inhibiting the formation of $ZnSO_4$, the mixt is leached to remove such sulfates as have been rendered water sol., and the sludge is subjected to a "hydrosepn action" to remove undissolved $CaSO_4$ particles held in suspension

Roasting sulfide ores such as zinc blende. METALLGES A.-G. Brit. 335,852, June 4, 1929 See Ger 502,586 (C. A. 25, 470)

Agglomerating pulverulent hematites. BAAG CHRISTIANSEN. U. S. 1,792,413, Feb. 10 The pulverulent ore is crushed to such fineness that the main portion of the mass has a grain size of 0.04 mm. or less but not a colloidal or slimy consistency, and the material thus prepd is heated

Coked agglomerates. EARL H. BUNCE (to The New Jersey Zinc Co.) Can. 308,891, Feb. 24, 1931. An intimate mixt. of zinciferous material and carbonaceous material with a moisture content of 6-12% of the dry wt. is initially heated by direct contact with a gaseous heated medium passed therethrough of sufficient vol. and high temp. to form promptly a shell of coke on each agglomerate, and the heating is continued until coking is completed

Coked agglomerates. GEORGE T. MAHLER and EDWIN C. HANDWERK (to The New Jersey Zinc Co.) Can 308,894, Feb. 24, 1931. A mixt. of zinciferous material and carbonaceous material is subjected to an initial temp. not exceeding 550° and the continued application of heat for 2-4 hrs. at temps. gradually increasing to the coking temp., whereby the agglomerates first undergo a drying stage without softening and next become plastic and finally become coked throughout. The agglomerates possess satisfactory residue strength for vertical retort smelting

Casting metals under pressure. P. HASSLER. Brit. 336,129, Nov. 7, 1929 App. and details of operation are described.

Continuous casting of metals between rolls, etc. H. HARRIS. Brit. 336,727, Aug. 27, 1929 Mech. features.

Casting metals with sheathings of other metals. I. SUGIMURA, K. MAGARISAWA and M. YASUMOTO. Brit. 337,284, Jan. 8, 1930. In casting a metal sheathing on a core metal having a higher m. p., lower sp. gr. and greater surface tension than the sheath metal, when the 2 metals are non-alloying with each other (as in the case of cast iron or iron alloys with Cu, Cu alloys, Pb or Sn or of Zn or Al with Pb), the melted metals are mixed in a vessel in which they assume such a configuration that when poured into a mold the core metal is suitably covered by the sheath metal. Various details of use of reducing agents, fluxes, etc., are described

Trough for centrifugal casting machines with horizontal rotary molds. AURELIO POSSENTI and CARLO SCORZA. Ger. 516,009, July 28, 1929

Apparatus for centrifugal casting of articles such as metal pipe. FRANK G. CARLINGTON (to Ferric Engineering Co.) U. S. 1,702,069, Feb. 10. Structural features.

Centrifugal casting of metal pipes, etc. FERRIC ENGINEERING CO. Brit. 337,105, Aug. 22, 1929 App. and various details of operation are described

Casting metal pipes in centrifugal molds. L. A. CAMEROTA. Brit. 336,857, Dec. 5, 1929 Mech. features

Mold for casting hollow metal bodies. MANNESMANNRÖHREN-WERKE. Ger. 513,894, Nov. 5, 1927 Details are given.

Water-cooled molds. ANTON MONT. Ger. 513,930, Oct. 27, 1927 The molds are for casting roller plates, axles, rods and tubes

Casting aluminum and its alloys. VEREINIGTE ALUMINIUM WERKE A.-G. (Hans Bohner, inventor) Ger. 516,129, Sept. 7, 1929 Local variations in the compn. of the cast metal are compensated, and castings of more uniform compn. are obtained by feeding the casting with an alloy differing in compn. from the casting. Thus, an ingot may be cast from an Al alloy contg. 5.95% of Cu, and the casting fed with an alloy contg. 7.26% of Cu

Furnace for reheating ingots. OERNBAU-GEZ. M. B. H. Fr. 694,653, April 28, 1930

Furnace suitable for the heat treatment of sheet steel, etc. RALF S. COCHRAN (to Surface Combustion Co.). U S 1,792,074, Feb. 10.

Metallurgical furnace suitable for steel treatment and decarbonization. FRED H. LOFTUS. U. S. 1,792,021, Feb. 10.

Continuous-process furnace suitable for carburizing metal articles such as those of steel. CHARLES T. WILLARD and RICHARD KAISA (to Singer Mfg Co.). U. S. 1,792,456, Feb. 10.

Tube or muffle furnace for heat treatment of wire. H. B. BLYTHE. Brit. 336,947, June 22, 1929.

Heat-treating furnace for enamelled articles. JOHN C. CROMWELL (to Steel Sanitary Co.). U S 1,792,284, Feb. 10.

Gas-heated doorless hardening furnace for tools. ERICH REIMANN. Ger. 513,601, June 15, 1928.

Apparatus for removing the products from the bottoms of blast furnaces. KOHLENSCHIEDUNGS G M B H. Ger. 513,611, June 23, 1927.

Apparatus for charging Scotch-hearth furnace such as those used for smelting ores. HUGH R. MACMICHAEL (to American Smelting and Refining Co.). U. S. 1,791,677, Feb. 10. Structural features.

Bessemerization of mats. THE INTERNATIONAL NICKEL CO. Fr. 694,092, April 17, 1930. Mats of Ni or alloys contg Ni such as Monel metal are bessemerized to remove S and reduce the mats by melting the mat in a converter and introducing a mixt of air and superheated steam. The mat is blown at a temp. of 1100-1500°, then carried to a temp. of 1500-1600° while continuing the blowing, the blowing mixt. being then almost completely free from air or O.

Reducing metallic compounds. GUSTAF N. KIRSEBOM. Fr. 694,283, April 23, 1930. Metallic compds are reduced by the reaction of As on a molten alkali metal compd. such as NaOH, in the presence of the metallic compds to be reduced. Pb, Cd, Sb and Bi may be obtained in this way from their oxides.

Extrusion manufacture of metal tubes, etc. P. MULTHAUPT. Brit. 336,688, July 20, 1929. In effecting the extrusion of liquid or plastic iron, steel or other metal, the matrix, mandrel, dies, etc., are formed of W or of an Fe-free alloy of W with C, Si, Zr and "metals analogous to W." Scale is removed by heated or high pressure gases.

Apparatus for cutting metals by fusion. L. J. HANCOCK. Brit. 336,433, Oct. 15, 1929. Structural features.

Apparatus for washing thin metal sheets. FRITZ GRAH. Ger. 516,208, Jan. 11, 1929.

Finely divided metals from carbonyls. I. G. FARDENING A. G. Brit. 336,007, July 10, 1929. See Fr. 690,991 (C. A. 25, 1211).

Metal acetylides. DEUTSCHE GOLD-UND SILBER-SCHNEIDANSTALT VORM. ROESSLER. Brit. 336,516, Feb. 7, 1929. Acetylides such as those of the alkali metals or their alloys with each other are made by continuously exposing fresh surfaces of the molten metals or alloys to the action of C_2H_2 (suitably in a rotating tube from which the air has initially been displaced by N). For making mono-Na acetylide, 60 hrs. are required at a temp. of about 170-80°, for di-Na acetylide, a temp. of 210-20° is used.

Aluminum salts. E. HEEZFELD and H. WALKER (to Colloid-chemische Forschungs A.-G.). Brit. 336,181, Jan. 7, 1929. Fe-free Al salts such as the sulfate are obtained from Fe-bearing solns. by pptn with alc. The soln and pptn may be repeated.

Dezincing lead. JESSE O. BETTEGATON (to American Smelting and Refining Co.). U. S. 1,792,210, Feb. 10. Precooled Cl gas is confined in a reaction tank, and molten lead to be dezincized is continuously recirculated through the Cl while maintaining the lower end of the tank sealed by the metal of the bath outside the tank to prevent escape of the Cl. App. is described.

Tin. WALTER PORTSCH and ROMAN VON ZELKOWSKI. Ger. 513,697, Aug. 15, 1928. Sn is obtained from raw material by beating it to 250° with NaSH. Ag tech NaSH-soln may be employed. In the example, zinc blende contg. about 10% Sn is heated with a 25% soln. of NaSH.

Tungsten. A. PACZ. Brit. 337,160, Oct. 3, 1929. In forming W ingots by compacting W powder and sintering, freedom from residual O is insured by mixing the powder with a more volatile metal such as Ag, compacting, sintering and volatilizing the added metal such as Ag to form a pure unworked W ingot suitable for use in the manuf. of filaments, spark points, elec. contacts, etc. Various methods are described involving admixt. of W and Ag compds, with or without a Th compd, followed by reduction and by volatilization of the Ag. Cf. C. A. 24, 5280.

Finely divided metals such as combined zinc and copper. WM. C. WILSON. U. S.

1,792,262, Feb 10 A metal such as Cu is reduced from one of its salts such as CuSO_4 by effecting intimate contact of the salt in a dispersed state with a precipitant metal (such as Zn in the case of CuSO_4), and the particle size of the liberated metal is controlled by the degree of comminution (less than 40 μ) of the precipitant metal

Magnetic testing system for testing materials such as sheet iron or steel, electric coils, etc. JAMES T. SERDUKE (to General Elec. Co.) U. S. 1,792,249, Feb 10

Rolling sheet, hoop or band iron. FRANZ JORDAN U. S. 1,792,377, Feb 10 Various mech. details are described for rolling an iron core between covering sheets of Al so thin that under the high roll pressures employed the Al is driven substantially completely into the pores of the iron core and forms substantially no exterior plating on the iron

Iron and steel. If HAGEMANN Brit. 336,054, July 18, 1929 In producing a non-siliconized iron or steel which can be easily welded, the C content of the bath is reduced (suitably to about 0.02%), and there is afterward added, in the furnace or ladle, an Al-Si alloy of such compn that it readily slags off. Mn low in C may be added to facilitate working of the product, and the latter may contain C 0.02, Mn 0.5, Cu 0.09, P 0.02 and S 0.003%. The Al-Si alloy added may contain Al 30, Si 50, Fe 19 and C 0.25%, or an alloy may be used contg Al 18-22 and Si 40-50%. Cf. C. A. 24, 2102

Steel. FOLDHÖTTE Austrian 120,608, May 15, 1927. Steel resistant to corrosion and capable of being forged contains C 0.2-1, Ni 20-25, Cr 12-20, Mn less than 0.5 and Si less than 0.5%, specifically C 0.5, Ni 22, Cr 15, Mn 0.4 and Si 0.3%

Steel. ANDRÉ P. MAHOUX Fr. 695,015, Aug. 8, 1929 Steels are hardened superficially by creating an oscillating field between the metallic piece and the piece to be treated, the dielectric being composed of a carburized or N atm.

"Stainless steel," etc. W. P. MARTIN and J. A. DEALYN Brit. 336,024, July 22, 1929 "Stainless steel" and similar Fe alloys contg. high proportions of Cr or Ni (or mixts. of metals adapted to form such alloys) are melted in a crucible in a furnace, the temp. of which is increased gradually and substantially uniformly to the casting temp. (especially at the temp. range between 1000° and the casting temp., which may be 1800-2400°). The heating may be at the rate of 100-400° per 15 min. Deoxidants such as Al may be added, together with other materials such as Cr hydrate if desired, just before casting, and the metal may be cast into molds made of 2 parts molding sand and 1 part sea sand bonded with molasses, the cores being formed of molding sand 1 and sea sand 2 parts

Manganese steels. TAYLOR WHARTON IRON & STEEL CO. Brit. 336,091, Oct. 4, 1929 Steels contg. Mn 10-15% and C 0.3-0.85% also contain up to 4% Ni and less than 0.4% Si, with or without 8% or less of Cr, and may be toughened by heating to 800-1150° and cooling in air

Case-nitriding of steel. AUGUSTUS B. KINZEL (to The Electro Metallurgical Co. of Can., Ltd.) Can. 309,049, Mar. 3, 1931 An article of ferrous material is heated to 700° for 2 hrs. in the presence of H_2 , then heated in the presence of V to 800° for 2 hrs., and then heated in the presence of NH_3 to 500-580° for a time sufficient to make a hard surface layer on the article

Case-hardening by nitridation. AUGUSTUS B. KINZEL (to The Electro Metallurgical Co. of Can., Ltd.) Can. 309,047, Mar. 3, 1931 Ferrous material is heated with Al to about 550° to form an aluminized coating and then subjected to a temp. of about 510° in the presence of NH_3

Case-hardening by nitridation. AUGUSTUS B. KINZEL (to The Electro Metallurgical Co. of Can., Ltd.) Can. 309,018, Mar. 3, 1931 Fe and steel articles (contg. alloying elements as Mn, Si, Cr and particularly V and Al) are case hardened by heating in a molten nitriding bath contg. one or more metal cyanides (as mixts. of NaCN and KCN, or of NaCN and $\text{Zn}(\text{CN})_2$, or of KCN and $\text{Mg}(\text{CN})_2$) at a temp. below 580°.

Alloys. W. C. HERABUS G. M. B. H. Ger. 516,490, May 20, 1930 Addn to 437,173 Ger. 437,173 describes alloys for making pen nibs comprising Ru 40-60, Os 35-50 and Pt 5-15%. There are now used alloys contg. Ru 60-80, Os 10-35 and another metal of the Pt group 5-25%

Alloys. THEODOR KIRTL. Austrian 120,383 July 15, 1930 In fractionally crystg. fused alloys so as to obtain alloys of different compn., the melt is slowly circulated during the crystn., with the result that larger crystals are obtained. Circulation may be effected by elec. induction heating

Alloys. WALTHER MATHESSUS and MARTIN W. NIEFELD Ger. 513,623, April 27, 1929 Addn to 441,071 Pb is hardened by adding alkali or alk. earth metal, Mg or Al and 0.1% Bi, Cu or Sn. The Pb alloy may contain 0.6 to 0.65% Na, 0.75

to 0.1% Ca, 0.2 to 0.25% Mg and 0.10% Al, the ratio Ca:Mg:Al being kept at the value 8:2:1.

Alloys. KARL SCHMIDT G M B H Ger 516,200, Aug 27, 1926. Al alloys for making pistons contain Si 29-35 and Cu up to 5, with or without Mg and (or) Mo up to 5%. Cf C A 24, 3982.

Alloys. GEORG WACK Ger 513,513, Sept 7, 1929. Objects resistant to HCl and H₂SO₄ are made from alloys contg NiSb 0.5 to 70, Ni 0 to 30 and W or Mo 0.5 to 60%. The NiSb may be replaced wholly or partly by W.

Alloys. THE INTERNATIONAL NICKEL CO, INC Fr 691,020, April 12, 1930. Alloys which are resistant to corrosion contain Fe 20-80, Ni 12-79, C 0.01-2 and Cu 1-32%. An example contains Cu 6, Ni 34, 1 e 10 and C 0.2%. Fr 694,930 describes an alloy contg Cu 1-20, Ni 12-30, C 0.01-2% and Fe the rest. An example contains Ni 25.5, Cu 4.5, Mn 0.6, C 0.1% and Fe the rest. Cf C A 25, 677.

Alloys. SOC DES TOURNILLES DE ST LOUIS Fr 695,050, Aug 9, 1929. An alloy for protecting Fe, etc., against corrosion contains Cd and Sn, Cu and Pb or Sb being optional addns. An example contains Cd 10, Cu 10, Sn 60 and Pb 20%.

Alloy for bearing surfaces. H C HALL Brit 336,881, Dec. 30, 1929. "Commercial Zn" is alloyed with Al 8-18 Cu and Ni up to 1% each and preferably only traces of other metals. The Zn is added after melting the other metals together.

Alloy for jewelry. VICTOR WULF Can 208,831, Feb 24, 1931. To 5 lbs 10 oz of melted Cu are successively added 3 oz fuller's earth, 7 oz Na₂B₄O₇, 3 oz ammoniated mercury, 12 oz Sn, 3 oz MgO and 1 oz alc., and the mixt is agitated and boiled. The resultant alloy simulates Au and is malleable ductile, immune from tarnishing and suited for jewelry.

Alloys for sealing to glass such as in vacuum tube and incandescent lamp manufacture. N-V PHILIPS' GLORILAMPFABRIK BV Brit 337,089, Aug 6, 1929. Ring- or disk-shaped members of metal to be sealed to material such as glass are formed of an alloy contg Ni (preferably over 60%) together with other metals such as W, Mo, Ta, V, Nb, Co or Si, the coeff of expansion of the alloy being at least approx. equal to that of the insulating material with which the alloy is to be sealed. One or more substances such as Cr, Al or Mg which facilitate the formation of a coherent film of oxide may be added to the alloy in proportion preferably less than 5%.

Magnetic alloys. W. S. SMITH, H. J. GARNETT and W. F. RANDALL. Brit. 336,048, June 22, 1929. Alloys substantially free from C and composed mainly of Fe and Ni are given a high and const. permeability by heating to 600° or over until completely annealed, cooling and further heating to not over 700° but to at least 500° above the magnetic change point of the alloy (and in the case of alloys contg not more than 60% Ni preferably at least 100° above the change point). The alloy is preferably subjected to mech. straining between the 2 heat treatments, and the sp. resistance of the alloy may be increased by the addn of up to 10% of one or more of the elements Mn, Cr, Cu, Mo, W, V, Al or Si. Various details and examples are given. Cf C A 24, 1074.

Aluminum alloys. H. T. TILLQUIST and J. HLABEN Brit. 337,099, Aug 17, 1929. See Fr 680,959 (C A 24, 3983).

Alloys of copper. P. M. G. METAL TAUSTR, LTD Fr 600,875, July 26, 1929. Cu alloys are made by adding to Cu or a Cu alloy a hardening alloy contg Fe, Si and Cu. The Fe constituent of the hardening alloy may be obtained by alloying steel and cast Fe, and the steel may include one or more of the metals Ni, Cr, Mo, W and V. The hardening alloy may also contain Mn or P or both. Examples are given. Cf C A 24, 5712.

Copper alloys. VICKERS-ARMSTRONGS, LTD, W. MACHIN and W. B. O'B. GOURDIN-LOCK Brit 335,950, April 4, 1929. Alloys such as those of high Cu content are made by adding to Cu or a Cu alloy a hardening alloy comprising Fe, Si and Cu (which may be prep'd in part from steel contg also Si, Cr, Mo, W or V), and Mn and P also may be present in the hardening alloy. Various details are given, and a final alloy may comprise Cu 88, Zn 2 and hardening alloy 10%. The alloys may be heated to over 800° and quickly cooled in water or oil, in some cases being reheated to not over 450°.

Copper alloys containing tin and zinc. MICHAEL G. CORSON (15% to George H. Corey) U. S. 1,792,146, Feb 10. Alloys of Cu contg also Sn 4-18 and Zn 10-35% are treated, to improve their properties as lubricated wearing surfaces, by heating for a proper time and at a suitable temp. to bring about a uniform β state of the alloy constituents, then quickly cooling to produce a hard acicular structure, reheating to a temp. above the eutectoidal transformation point for the alloy to produce a structure characterized by the presence of β crystals uniformly distributed throughout a matrix of α crystals, and then cooling to effect stabilization.

Magnesium alloys. I. G. FARBERVIND, A.-G. Brit. 336,019, April 25, 1929. Products such as high percentage Mg alloy castings are improved by heating at a temperature in which the solubility of the alloying component or components is increased and then subjecting them to cooling which is considerably retarded artificially especially in the range 300–100°. Details are given of the treatment of Mg-Al alloys containing 7.5% or more Al, Mg-Zn alloys containing 1.8% or more Zn, and Mg-Pb alloys containing 18% or more Pb. An alloy containing 8.2% Al may be heated to 410° for 72 hrs., then transferred to an annealing furnace and cooled down to 100° in 24 hrs. (the cooling from 300° to 100° taking 18 hrs.)

Magnesium and its alloys. I. G. FARBERVIND, A.-G. Brit. 336,498, Jan. 2, 1929. Mg and Mg alloys are refined by treating the molten metal with a material such as Mn, Cr, Mo or Si which on cooling combines with impurities such as Fe for separation from the remaining melt. A small proportion of the added reagent such as Mn may remain in the product. Cf. C. A. 25, 678.

Preventing corrosion of aluminum. G. ADOLPH. Brit. 336,854, Nov. 29, 1929. To prevent attack of Al by alk. solns., such as alk. bleaching liquors, the latter are admixed with an alkali silicate, H_2O_2 compounds such as Na perborate or percarbonate, or mixtures of these.

Rust prevention. KARL DARVES. Fr. 693,905, April 14, 1930. The formation of white rust on galvanized objects or other metallic surfaces is prevented by exposing the surfaces to the action of free air or the gases produced by the combustion of S before storing in closed spaces.

Rustproofing iron. W. BÖHMES. Brit. 335,868, Jan. 27, 1930. The cleaned surfaces are treated either with (1) a mixture of PbO_2 , PbO , glycerol and waterglass, or (2) a mixture of zinc white and waterglass, or (3) a ZnCl₂ solution. Specific formulas are given.

Rustproofing iron and steel by coating with phosphates. VAN M. DARSY (to Parker Rust-Proof Co.) U. S. 1,791,715, Feb. 10, 1930. An iron or steel surface carrying a phosphate coating containing a metal such as Cu having a potential different from that of iron or steel is treated with a solution of a material containing the chromic acid radical such as K chromate.

Welding aluminum-steel. H. J. SCHIFFLER. Brit. 336,195, March 25, 1929. Scale-resisting steels containing Al (and which may also contain other components such as Cr and Mo) are welded by electrodes or rods of scale-resisting steel alloys not containing Al (but which may contain Cr, Ni and Mo) in order to produce a scale-resisting weld. Various formulas of alloys used are given.

Welding bronze. L. J. THRENNHAM. Brit. 336,402, Oct. 1, 1929. A welding rod for use in welding bronze by the blow-pipe flame is formed of bronze with a coating such as H_2BO_3 49, Na borate 9, metal oxides such as those of Fe and Cu 30 and Na silicate 12 parts.

Welding magnesium and its alloys. I. G. FARBERVIND, A.-G. Fr. 694,137, April 18, 1930. See Ger. 508,352 (C. A. 25, 680).

Electrodes for spot-welding copper. P. KRUYP, A.-G. Brit. 335,876, March 26, 1929. Electrodes are used containing one or more carbides of W or similar high melting point metal, and they may also contain a small quantity of metals such as Fe, Co or Ni of lower melting point, e.g., 6% of Co. The material sold under the name "Widia" is suitable.

Weld iron. ARTURO DOSSMANN. Ger. 513,732, April 27, 1928. Small iron waste is used as starting material and is prevented from oxidation by a coating of colophony or molasses before mixing with the wood charcoal or soot and limestone.

10—ORGANIC CHEMISTRY

CHAS. A. ROUELLER AND CLARENCE J. WEST

Catalytic addition of gaseous hydrochloric acid to unsaturated hydrocarbons. W. J. PROYROWSKI AND J. WINKLER. *Przemysł Chem.* 15, 25–36 (1931).—The 20–40° fraction of cracked benzene was used as the starting point for this synthesis. It was characterized by n_D^{20} 1.532, amylene content 31% by wt. Its complete distillation curve at 1° intervals is given. Two hundred cc. portions of this pentane amylene fraction with slight excess of HCl were passed over catalysts through a glass apparatus for 4 hrs. The yield was figured on the basis of change in d_4^{20} and distillation fractionation. It was found advantageous to pass the raw materials over the catalysts held in 2 tubes, the 1st of which was kept at 70° and the 2nd at 150°, in which case a predominance of secondary chlorides was ob-

tailed. If the sequence of these temps. was reversed a mixt. of all possible chlorides resulted. The presence of moisture interfered with the formation of secondary chlorides, and its absence never had any detrimental effects. The catalysts here studied were the chlorides of bivalent, trivalent and quadrivalent metals mounted on activated charcoal and dried at 150° for several hrs. to const. wt. The catalysts themselves caused no polymerization in the benzene in the absence of HCl. The reaction between HCl and the hydrocarbons was vanishingly small without the catalysts, it was marked in the presence of activated charcoal. The catalysts are classified in 3 groups: (1) CrCl_3 and FeCl_3 , which are inactive, (2) HgCl_2 , CuCl_2 , CdCl_2 , MnCl_2 , BiCl_3 , AlCl_3 and SbCl_3 , which are slightly active, (3) ZnCl_2 and SnCl_4 , both of which are strongly active, SnCl_4 having given a practically quant. formation of amylene monochlorides. The catalysts are subject to poisoning by S compds., which should be removed by activated charcoal previously partially poisoned. Sharp fractionation of the product showed it to be composed of Me_2CClEt , $\text{Me}_2\text{CHCHClMe}$, MeCHClPr and $\text{Me}_2\text{CCHClMe}$. A. C. Z.

Transformation of propylene into liquid hydrocarbons. A. MAILHE and R. RENAUDIE. *J. usines gaz* 54, 568-9. *Compt. rend.* 191, 851-4 (1930); cf. *C. A.* 24, 5283— C_3H_6 , heated over SiO_2 gel at 650°, furnishes a considerable proportion of liquid hydrocarbons, unsatd. and aromatic, C_6H_6 and PhMe predominate. Solid hydrocarbons, $\text{C}_{11}\text{H}_{22}$, $\text{C}_{14}\text{H}_{30}$, etc., were absent, their formation beginning at 700-20°. Traces of Fe led to excessive formation of C and inactivation of the catalyst. K. H. ENGEL.

Synthesis of squalene. P. KARRER and A. HELFENSTEIN. *Helv. Chim. Acta* 14, 78-83 (1931).—In a recent paper Karrer (*C. A.* 25, 519) suggests the sym. formula ($\text{Me}_2\text{C} \cdot \text{CHCH}_2\text{CH}_2\text{CMe} \cdot \text{CHCH}_2\text{CH}_2\text{CMe} \cdot \text{CHCH}_2$), (I) for squalene rather than those proposed by Heilbron (*C. A.* 23, 3702). Such a structure may be considered as made up of isoprene residues and in this way looks plausible in the light of the Karrer proposed structure for lycopene and carotene. On the assumption of the above structure K. and H. synthesized I by treating farnesyl bromide, $\text{Me}_2\text{C} \cdot \text{CHCH}_2\text{CH}_2\text{CMe} \cdot \text{CHCH}_2\text{CH}_2\text{CMe} \cdot \text{CHCH}_2\text{Br}$ (II) with K or Mg, whereby 2 mols. of II were joined through the elimination of Br from each. The synthetic I forms a mixt. of 2 isomeric hexahydrochlorides with HCl as does I from natural sources. These hydrochlorides may be sep'd by hot acetone into a hexahydrochloride m. 143-5° (0.25-0.20 part) and one m. 108-10°, exactly as may be done when they are prep'd. from natural I. Crystals of the hexahydrochloride m. 143-5° from synthetic source were compared with those from a natural source and in both cases rhombic forms, chiefly *m*(110) and *c*(001) and seldom *b*(010), were observed. The angle (110):(110) is about 59° so that $a:b = 0.56:1$ in both. Optical orientation $n_{\parallel} || b$, $n_{\perp} || a$, $n_{\parallel} || c$; optical character, neg. X-ray absorption of the powder of both preps. gave 11 lines with Fe-K radiation which agree in position and intensity. The yields of I were poor (about 5%) because K. and H. could not prep. pure II. Probably the difficulty is due to the known tendency of halogen in similar compds. to migrate to a tertiary C atom or the reverse: $-\text{MeC} \cdot \text{CHCH}_2\text{Br} \rightleftharpoons -\text{MeCBrCH} \cdot \text{CH}_2$. In agreement with this assumption it was possible to prep. from nerolidol, $\text{Me}_2\text{C} \cdot \text{CHCH}_2\text{CH}_2\text{CMe} \cdot \text{CHCH}_2\text{CH}_2\text{CMe}(\text{OH})\text{CH} \cdot \text{CH}_2$ (III), a mixt. of Br compds. which yielded I on treatment with K or Mg. K. and H. regard I as a possible mother substance of cholesterol (cf. André and Canal, *C. A.* 23, 3364). Ten g. of farnesol (IV) or III (3 mol.) was added dropwise to a mixt. of 1 g. dry pyridine (1/2 mol.) and 5 g. PBr_3 (1.2 mols.) at -5° and allowed to stand overnight at room temp., washed with ice-cold NaHCO_3 soln. and extd. with ether. The ether soln. was washed successively with dil. H_2SO_4 , NaHCO_3 , and water, then dried and distd. A fraction b.p. 110-20° with 8.1% Br and one b.p. 125-300° (about 9 g.), with 17.3% Br were collected. Redistn. caused decompn. and lowering of the Br content. This impure bromide (V) was used although the theoretical percentage of Br is 23.07. Other methods of bromination or chlorination of III or IV gave poorer yields of less pure material. Into 9 g. crude V at 100° was dropped 3 g. K. and the temp. raised to 135° for 0.5 hr.; the cold product was treated with abs. ether and excess K. removed by filtration through glass wool. The soln. was washed with dil. H_2SO_4 , and then with water, dried and distd. The fractions sep'd were one of 3.5 g., b.p. 100-20°, the squalene fraction of 2 g., b.p. 210-40° and a residue of 1.5 g. which yielded no squalene hexahydrochloride (VI) with HCl. The 2 g. squalene fraction yielded 0.05-0.15 g. of VI on treatment with HCl. Impure V (15 g.) was treated with 1.7 g. Mg and yielded 3 g. impure I, which on conversion to VI gave 0.55 g. Impure V (28 g.) treated with 2.7 g. Mg yielded 6.8 g. I and this gave 1.3 g. VI. One g. squalene, from natural sources gave 0.5 g. VI; if the Br content of V is taken as a basis of its purity the yield of I from pure V is about 20%.

ODEN E. SHEPPARD

Catalysis of polymerization by ozonides. RAY C. HOUTZ and HOMER ARKINS. *J. Am. Chem. Soc.* 53, 1038-63 (1931).—A mixt. of the ozonides of 2 octylenes has

been found to catalyze the polymerization of styrene, indene, $(\text{C}_6\text{H}_5\text{CMe})_2$, a methylpentadiene and furfuryl alc. It has been found to be inactive as a polymerization catalyst toward stilbene, $\alpha\text{-HOC}_6\text{H}_4\text{CH}_2\text{OH}$, $\text{Me}_2\text{C}=\text{CHMe}$ and the mixt of octylenes from which it is made. The catalyst was active at room temp and 100° and in a soln in EtMe . The ozonide underwent a change on standing, especially during the first 2 days after its prepn, and lost a large part of its catalytic power. C. J. WEST

The halogenation of the paraffin hydrocarbons. GUSTAV EOLOFF, R. E. SCHAAPE AND C. D. LOWRY, JR. *Chem. Reviews* 8, 1-80 (1931).—A summary of the work which has been done on the halogenation, particularly the chlorination and bromination, of paraffin hydrocarbons. An extensive bibliography is given. Problems for research are suggested. LOUISE KELLEY

Preparation of alkyl iodides. HAROLD S. KING. *Proc. Trans. Nova Scotia Inst. Sci.* 17, 242-3 (1930).—In the prepn of alkyl iodides by a modification of Adam's method the use of anhyd. alc. is not necessary (*Proc. Trans. Nova Scotia Inst. Sci.* 16, 87 (1924)). The addn of 10% by wt of water to MeOH did not reduce the yield of MeI . A MeOH soln of I is slowly run into the reaction flask contg. P and more MeOH and heated at $70-6^\circ$. The refluxing liquid passes over I on its return to the flask. RACHEL BLOOM

Preparation of anhydrous alcohol, and the detection and determination of water in alcohol. P. ADICKES (WITH W. BAUMHEAT AND O. LUCKE). *Ber.* 63B, 2753-60 (1930).—When HCO_2Et is added to alc. NaOEt it is almost instantly saponified in the cold by the slightest traces of water in the alc. according to the equation $\text{HCO}_2\text{Et} + \text{NaOEt} + \text{H}_2\text{O} = 2\text{HOEt} + \text{HCO}_2\text{Na}$, with pptn. of the very difficultly sol. HCO_2Na . As HCO_2Na can be detd. easily and accurately the reaction can be used for the 3 purposes indicated in the title. After this work had been started, it was found that Smith had made use of the same principle, employing $(\text{CH}_3\text{CO}_2\text{Et})_2$ instead of HCO_2Et (*C. A.* 21, 3428). $(\text{CH}_3\text{CO}_2\text{Et})_2$ is much more difficult to saponify than HCO_2Et and the anhyd. alc. he prepd. by its means contained 0.01 g. of the ester per l. With HCO_2Et , the equl. in the catalytic decompn. of HCO_2Et by alc. NaOEt ($\text{HCO}_2\text{Et} = \text{CO} + \text{HOEt}$) lies practically completely on the CO side at the l. p. of the alc., and by using an excess of NaOEt , refluxing until equl. is established and distg. away a first fraction, an alc. can be obtained whose ester content is at least 10 times less than that of the alc. obtained with $(\text{CH}_3\text{CO}_2\text{Et})_2$. When neutral HCO_2Na is boiled a long time with alc. and the alc. is then slowly distd. off, the residue becomes slightly alk. because of loss of HCO_2H as HCO_2Et ; the alk. corresponds to 0.0037% ester in the alc. distd. off, or a water content of only 0.0009%. The process for prepg. anhyd. alc. consists simply in dissolving Na in the alc., adding HCO_2Et , heating under a reflux until no more CO is evolved and then adjusting the cooling water in the reflux so that the alc. distils off. The alc. so obtained contains 0.03% water and 0.0015% HCO_2Et . The soly of HCO_2Na in alc. in the presence of NaOEt is so small at 0° (0.05 in 100 g. alc.) that by addn. of NaOEt soln. and HCO_2Et as little as 0.013% water in alc. can be detected qualitatively by the turbidity produced. For quant. detns., the HCO_2Na can simply be filtered off when the quantity of water is large (2.5-3%), with smaller quantities of water the alc. and undecomposed ester must be distd. off (finally *in vacuo*); with very small quantities of water (0.1%) there must be applied a correction factor for the water or acid content of the HCO_2Et , detd. by running parallel detns. with different quantities of HCO_2Et . The HCO_2H was detd. by weighing the HgCl_2 pptd. from HgCl_2 ; the detn. cannot be carried out in alc. since the latter also reduces HgCl_2 , although but slowly. With the larger quantities of water the detn. can be made volumetrically by filtering off the HCO_2Na and titrating it with an oxidizing agent. The detn. can also be made by using a known quantity of HCO_2Et and measuring the CO evolved by decompn. of the excess, in this case allowance must be made for the ester vapors carried along by the CO and for the water and acid content of the HCO_2Et , and the method is accurate enough only for water contents of 1% or more. C. A. R.

Preparation of absolute ethyl alcohol. E. GORDON YOUNG. *Proc. Trans. Nova Scotia Inst. Sci.* 17, 248-55 (1930).—A comparative study has been made, with reference to % yield and cost per l. of product, of 3 methods of prepg. abs. alc. for ordinary chem. procedures. By the carbide method (1) 100 g. 95% alc. was refluxed 2 hrs. with 25 g. CaC_2 and 1 hr. further with 1 g. CuSO_4 . A 2nd treatment with CuSO_4 was necessary to remove odor. The yield was 95% of 99.7% alc. By the CaH_2 method (2) 2000 cc. alc. and 1500 cc. C_2H_5 were distd. in a fractionating column with 22 bulbs, 300 cc. abs. alc. was obtained. By the CaO method (3) alc. was refluxed with varying amts. of CaO for varying periods of time. The time required was found to vary inversely with the amt. of CaO used. 11 95% alc. refluxed with 300 g. fresh CaO for 4 hrs. gave an

80% yield. Conclusion: method 1 gives the best yield at a cost slightly higher than 3; method 2 is impracticable on a lab. scale, method 3 is the cheapest. R. B.

Organic compounds of thallium. R. C. MENZIES. *Chem. News* 141, 305-6 (1930).—The Tl atom replaces the H of many OH compds., including simple and polyhydric alcs of the aliphatic series, with the formation of stable compds. which are generally less sol than those of the alkali metals. In many cases all the H of the OH of the polyhydric alcs and the H of the OH and the CO₂H of the HIO acids can be replaced, but well-defined Tl salts of the reducing sugars have not yet been prepd because of the oxidation of the sugars with the reduction of the metal. However, with compds in which the reducing groups are protected, well-defined Tl derivs are obtained, although with α - and β -Mie glucosides 1 of the 4 available H escapes replacement and with sucrose only 4 of the 8 available H are replaced. An explanation of these observations has not yet been found. C. J. PEDERSEN

Catalytic hydrogenation of esters to alcohols. HOMER ADKINS AND KARL FOLKERS. *J Am Chem Soc* 53, 1005-7 (1931).—Et laurate, myristate, valerate, cinnamate, trimethylacetate and succinate have been hydrogenated over a Cu chromite catalyst with the formation of the corresponding alcs in yields of 80-98%. The CO₂Et group in HOCH₂CM₂CO₂Et was hydrogenated to a CH₂OH group, and in addn the mol. underwent cleavage between the 2 and 3 C atoms. The H pressure during the hydrogenation approximated 220 atm and the temp was 250°. C. J. WEST

Esters of sulfurous acid. I. WALTER VOSS AND ERICH BLANKE. *Ann.* 485, 258-83 (1931), cf. C. A. 25, 69.—Dropping 1 mol. MeOH into 1 mol. SOCl₂ cooled in ice and salt and simultaneously passing through a rapid stream of CO₂ to stir the mixt. and remove HCl as formed, gives 50% of *Me chlorosulfinate*, MeOSOCl, b₁₀ 35°, d₄²⁰ 1.4186, slowly decomposing at room temp into MeCl and SO₂ and immediately decomposed by water. *Et ester*, b₁₀ 32°, b₁₀ 41-2°, d₄²⁰ 1.2826 (68% yield), *iso-Pr ester*, b₁₀ 34°, d₄²⁰ 1.2005 (47% yield). SOCl₂ and 2.2 mols MeOH with cooling gave 70-5% (MeO)₂SO, b₁₀ 126°, d₄²⁰ 1.2073, di-Et ester, b₁₀ 158° (85% yield); Pr ester, b₁₀ 192-25°, b₁₀ 80° (90% yield), boiling 24 hrs. causes little change; *iso-Pr ester*, b₁₀ 78°, b₁₀ 160.5-70.5° (33% yield); *Bu ester*, b₁₀ 114-5°, d₄²⁰ 0.9944, n_D²⁰ 1.430511, 1.43808, 1.44268 for α , β and γ , n_D²⁰ 1.43244 (94% yield); *iso-Am deriv*, b₁₀ 127-8°, d₄²⁰ 0.97293, n_D²⁰ 1.43547, 1.44315, 1.44820, 1.43757 for α , β , γ and D (72% yield); *cyclohexyl ester*, b₁₀ 182°, d₄²⁰ 1.0974, n_D²⁰ 1.48306, 1.49571, 1.4885 for α , β and D (54% yield); *di-Ph ester*, b₁₀ 143°, b₁₀ 152°, d₄²⁰ 1.2404, n_D²⁰ 1.56899, 1.58833, 1.57441 for α , β and D (34% yield). EtOSOCl and the calcd amt. of MeOH and C₂H₅N give *Me Et sulfite*, b₁₀ 53°, b₁₀ 141.5°, d₄²⁰ 1.1364, n_D²⁰ 1.41417, 1.42159, 1.42628, 1.41669 for α , β , γ and D. If C₂H₅N is not used, a mixt. of Me₂SO₂, Et₂SO₂ and Me₂SO₃ results. Me₂SO₂ (0.2 mol) and EtOH (1 mol.) with 0.1 cc. 12% EtOH-HCl as a catalyst, give 44% of Et₂SO₂. Similarly Et₂SO₂ and MeOH give Me₂SO₂. MeOS(O)₂Me results in 67% yield from MeSO₂H and MeI at 100°; b₁₀ 101-2°, b₁₀ 202-3°, d₄²⁰ 1.2975, n_D²⁰ 1.41384, 1.42027, 1.41576 for α , β and D. The kinetics of the hydrolysis of MeOS(O)₂Me and Me₂SO₂ by water, acid and alkali are discussed. The following values were obtained for pure water (H⁺ = 1) and (OH⁻ = 1): MeOS(O)₂Me, 2.98×10^{-4} , 0, 2.81×10^{-3} , Me₂SO₂, 0, 1.30×10^{-3} , 13.37; Et₂SO₂, 1.55×10^{-4} , 0, 9.5×10^{-3} , Me₂CO₂, 0, 0.54. The Na derivs. of PhOH, o-O₂NC₆H₄OH and β -C₁₀H₇OH with 1.1 mols dialkyl sulfite give 80% PhOMe, 88% β -C₁₀H₇OMe, 85% o-O₂NC₆H₄OMe, 88% PhOPr, 55% iso-AmOPr. With 3 mols Me₂SO₂ the o-, m- and p-C₆H₄(OH)₂ give 89, 76 and 80% of the di Me ethers, resp. PhOH and Me₂SO₂ with a few drops concd H₂SO₄ or EtOH-HCl give 9.3% PhOMe; β -C₁₀H₇OH gives 37% Me ether and 63% Et ether. One mol. PhNH₂ with the dialkyl sulfites gives the following yields of alkyl derivs: 96% PhNMe₂; 52% PhNEt, 31% PhNHET; 38% PhNPr, 44% PhNIPr; 15% PhN(iso-Am), 74% PhN(iso-Am). *Benzenesulfonethylanilide*, b₁₀ 183.5°, d₄²⁰ 1.1943, n_D²⁰ 1.57637; *p-toluenesulfonisoamylanilide*, m. 76-7°. PhNMe₂ and Me₂SO₂, heated at 135-40°, give 99% of trimethylphenylammonium methanesulfonate (I). (PhNMe₂)OS(O)₂Me, m. 188-9°; this also results from PhNMe₂ and MeOS(O)₂Me at room temp; on warming 24 hrs. with 12% KOH 53% of PhNMe₂ are split off; 5% MeONa splits off 96.5% in 5 hrs. PhNMe₂ and Me₂SO₂ on standing at room temp for several weeks deposited 10% of a yellow, very hygroscopic product, m. 120-1°, which may be (PhNMe₂)OSO₂OMe; this could not be changed into I by heating at 135°. *N-Methylpyridinium methanesulfonate*, red, very hygroscopic, m. 117-8°; C₅H₅N and Me₂SO₂ at room temp give the compd. C₅H₅N₂O₂S, m. 104-7°. The following yields of esters were obtained with

1.1 mols dialkyl sulfite and a drop of concd H_2SO_4 on warming 1-2 hrs: $BrOEt$ 96, $BrOBu$ 88, $p-BrC_6H_4CO_2Et$ 95, $PhCH_2CHCO_2Et$ 85, $MeCH_2CHCO_2Me$ 82, $(CH_3)_2CO_2Et$, 90, $CH_3(CH_2CO_2Et)_2$ 81; salol, 65, di-Et carbonates do not give esters under these conditions. II. Preparation of acetals and glucosides. W. Voss *Ibid* 283-93—Cyclohexanone (I) and Me_2SO_4 in $MeOH$, with $MeOH-HCl$ as a catalyst, gave 79% of the di-Me acetal, b_p 64°. In the same way $PrCHO$ gave 87% of the di-Me and 72% of the di-Et acetals, BzH gave 81% di-Me and 78% di-Et acetals; $p-MeOC_6H_4CHO$ gave 84% di-Me acetal and 1 gave 63% of the di-Et acetal. These new acetals will be described later. Ph_2CO and Me_2SO_4 gave 95% of di-Me acetal, m 106-7°. BzH (10.5 g), 12 g Me_2SO_4 and 27 g $BuOH$ (with $MeOH-HCl$) gave 80% of benzaldehyde di-Bu acetal, b_p 143-50°, d_4^{20} 0.9283, n_D^{20} 1.4790. The di-iso-Am acetal was obtained in 76% yield. The following yields of $Ph_2C(OMe)_2$ were obtained from 4.6 g Ph_2CO with (1) mol reagent in $MeOH$ with HCl as catalyst. Ph_2SO_4 1.0 g; $HC(OFt)_2$ 2.0 g; $Hf_2(OFt)_2$ 3.0 g; $Me_2C(OMe)_2$ 1.0 g; $Me_2C(OEt)_2$ 1.4 g. No acetal resulted with Me_2SO_4 in CH_2Cl_2 , Et_2O , etc., with HCl as a catalyst or in $MeOH$ with $AcOH$ or picric acid as catalyst. d-Glucose (36 g) with 25 g Me_2SO_4 and 100 cc. 3% $MeOH-HCl$ warmed about 2 hrs, gave 19 g α -Me and 2.8 g β -Me d-glucoside. Starch (250 g) gave 96% of α -Me glucoside. d-Galactose (36 g) formed 28.3 g of α -Me galactoside and only 0.1 g of the β -Me deriv. d-Mannose gave only the α -Me deriv. Details are given for the prepn. of Me rhamnoside, α -Me and α -Et l-arabinoside and α -Et d-galactoside.

C. J. West

Reaction between multivalent alcohols or phenols and arsenic compounds, especially arsonoacetic acid. III. BERTH. ENCLUND *J. prakt. Chem.* 129, 1-23 (1931); cf. C. A. 24, 1841.—By use of the previously described method the following values for L_2 have been detd. for the reactions of arsonoacetic acid (I) with $HO(CH_2)_nOH$ in $AcOH$ at 25° (values of n and L_2 given): 2, 0.295; 3, 0.069; 4, 0.052; 5, 0.060; 6, 0.068; 7, 0.047; 8, 0.057; 9, 0.043; 10, 0.045. With As_2O_3 in $BuOH$ contg. 5% H_2O values of L_2 were detd. as follows: $MeCH(OH)CH_2OH$, 0.128; $HO(CH_2)_2OH$, 0.012; *meso*-($MeCH(OH)_2$), 0.288, *di*-deriv., 0.240; pinacol, 0.340; $HO(CH_2)_3OH$, 0.013; $HO(CH_2)_4OH$, 0; $HO(CH_2)_5OH$, 0.11. With I in $AcOH$ the aldehydic diols gave the following values for L_2 , the 2 figures being for the *cis*- and *trans* forms: $Cal_2(OH)_2$, 0.375, 0.034; $Cal_2(OH)_2$, 0.420, 0.087; hydrindene-1,2-diol, 0.485, 0.050; tetrahydronaphthalene-1,2-diol, 0.475, 0.022. The diglycol esters of As_2O_3 are relatively weak acids, with excess of glycol there result complex triglycol derivs., thus $HAs(C_2H_4O)_3$ has been isolated as salts of the acid. While the glycol deriv. of As_2O_3 and I show a double mol wt. in $CHBr_3$, they are monomol. in Me_2CO . The pinacol deriv. (II) of I titrates as a monobasic acid in alc., using phenolphthalein as indicator, the other derivs. studied ($HIOAs_2G_2$, $HO_2CCH_2As_2G_2$ and $HOAsPn_2$, where G and Pn represent glycol and pinacol) are unstable and decompose in $EtOH$. Data are given for the hydrolysis of II by H_2O , this is complete in about 1 hr and in about half that time on the addition of a small amt. of HCl , it is fairly resistant to alkali. The elec. cond. is also reported for 3 concns., 2 at 25° and 1 at 0°. The final value agrees well with that for As_2O_3 . E. also discusses tensionless hetero rings.

C. J. West

Characterization of alkyl halides and organomagnesium halides. A. M. SCHWARTZ AND JOHN R. JOHNSON *J. Am. Chem. Soc.* 53, 1063-8 (1931).—The yields of $RMgX$ from a series of alkyl halides are reported, these $RMgX$ are characterized by transformation into the anilides of the corresponding acids with $PhNCO$, the cor. m. ps. for these anilides are given. The application of this method to certain branched-chain alcs. is limited by the fact that intramolecular rearrangements may lead to anilides that do not correspond in structure to the original alc. Thus, 3-methylbutan-2-ol and 2,2-dimethylpropanol both give the anilide of dimethylethylacetic acid, allyl alc. gives the anilide of crotonic acid. An optically active halide, *d*-2-bromooctane (in which the halogen is directly attached to the asym. C atom), led to a completely inactive anilide.

C. J. West

Partly abnormal reaction of β -substituted allyl bromides on organomagnesium bromides. CHARLES PRÉVOST AND JEAN DUCJAY *Bull. soc. chim.* [5], 47, 588-94 (1930).—While with organomagnesium bromides the simple allyl bromide reacts thus $RMgBr + CH_2=CHCH_2Br \rightarrow MgBr_2 + RCH_2CH=CH_2$ (Kirmann, C. A. 20, 2443), β -substituted allyl bromides give rise to a desmotropism in this reaction: $RCH=CHCH_2Br + R'MgBr \rightarrow MgBr_2 + RCH=CHCH_2R'$ or $MgBr_2 + RR'CH=CHCH_2$ (Prévost, *Ann. chim.* [10], 10, 121). The 3 reactions (1) $LiMgBr + EtCH=CHCH_2Br$, (2) $PhMgBr + EtCH=CHCH_2Br$ and (3) $EtMgBr + PhCH=CHCH_2Br$ were studied in detail and the 5 theoretically expected hydrocarbons, $EtCH=CHCH_2Et$ (I), $Et_2CHCH=CH_2$ (II), $EtCH=CHCH_2Ph$ (III), $PhCH=CHCH_2Et$ (IV) and $PhEtCH=CH_2$ (V).

were isolated. For the prepn of the β substituted allyl bromides (Bouis, C. A. 22, 213) the vinylcarbinols (Delaby, *Thesis*, 1923) were used. To the Mg reagent the allyl bromide in Et₂O is slowly added and the reaction is completed by several hrs. of refluxing. By using a slight excess of the Mg reagent immediately after decanting it from excess Mg, formation of brominated products as well as doubling of the allyl bromide is avoided. Although compds I-V are not new, the consts of II, IV and V were found to differ substantially from the literature. Reaction (1) I and II are sep'd by 4 distns. The total yield is above 80% and 3-4 times as much II as I is formed. II b 85°, n_D 1.3966, d_4 0.6918, mol refraction 33.94. Bromination yields the dibromide Et₂CHCHBrCH₂Br, b₁₁ 93.5°, n_D 1.5006, d_4 1.5251, mol refraction 49.81. KOEt forms isoamylacetylene, b 87°, n_D 1.4023, d_4 0.7272, mol refraction 32.24 (cf Lespiau, C. A. 24, 334). The Ag salt Et₂CHC CAg is very sol in EtOH. Reaction (2) III and V are sep'd by 7 vacuum distns, 2-3 times as much III is formed as V, and the total yield is excellent. V b₁₁ 71°, b₁₁ 191.5° (cor.), n_D 1.5030, d_4 0.8818, mol refraction 48.95. Reaction (3) IV and V are sep'd by 3 vacuum distns. In a total yield of 75-85% about twice as much IV is present as V. IV b₁ 87.5°, b₁₁ 91°, b₁₁ 217° (cor.), n_D 1.5318, n_D 1.5337, d_4 0.8921, d_4 0.8886, mol refraction 50.7, 51.0. Some *bis*cinnamyl was identified as a by-product. Data on I and III are also given. G. TONCHIES.

Highly polymerized compounds. XLVII. Morphology of highly molecular substances. II. Polyoxymethylene precipitates from solutions. HANS W. KORTSCHOTTER. *Ann* 484, 155-78 (1930), cf C. A. 25, 15—K ppts aq (CH₂O)_x solns (d_4 1.086) with H₂SO₄ and obtains 24 ppts., using various conditions of temp., amounts of H₂SO₄ and concns of (CH₂O)_x. Also, the effects of added substances (e. g., MeOH) and of the substitution of HCl and H₃PO₄ for the H₂SO₄ were studied. The significance of the ppts is discussed. C. J. WEST.

Relation of the structure of ketones to their reactivity and affinity in acetal formation. II. GEO. J. PREIFFER and HOMER ADKINS. *J. Am. Chem. Soc.* 53, 1043-8 (1931); cf C. A. 22, 383—HCO₂Et and Na react as follows: 3HCO₂Et + 4Na = 2CO + MeONa + 3EtONa (to be discussed in a later paper). The ratio of CO to HCO₂Et is independent of the amt of H(OEt), or of ketone acetal present in the mixt. subject to analysis; however, the ratio is dependent upon the amt. of EtOH and the amt. and kind of ketone present. A table is given showing the CO from HCO₂Et in the presence of various amts of EtOH and ketones. The reaction of the various ketones with HCO₂Et was carried out by mixing 0.1 mol of ketone and 0.1 mol HCO₂Et in EtOH to give a total vol. of 49 ml., then adding 1 ml. of EtOH contg. 0.0136 g. HCl; after 50-60 hrs. 5 ml. aliquots were removed and decomp'd with Na and EtONa; from the data the % acetal calcd. and equil. const., $K_g \times 10^4$, were det'd: Me₂CO, 04.8, 334; MeCOEt, 90.1, 0.83; MeCOCPh, CH₂Ph, 89.1, 0.66; MeCOPh, 80.2, 0.39; Me neopentyl ketone, 84.0, 28; (iso-Pr)₂CO, 65.4, 0.034; *tert*-BuCOMe, 50.1, 0.0101; *tert*-BuCOEt, 30.2, 0.0033; Ph₂CO, 33.7, 0.0026; *tert*-BuCOCHMe, 25.8, 0.0012; (*tert*-Bu)₂CO, 17.2, 0.0001. These results show that any substitution in the Me groups of Me₂CO results in a lowering of the extent of the reaction. The extent of conversion is decreased only slightly by the substitution of a PhCH₂ group for 1 of the H of Me₂CO. The effect of the replacement of a Me by Ph is marked, while 2 Ph groups reduce the % conversion to acetal to about 33% of the value for Me₂CO. The b. p. and d. for the various compds are given. C. J. WEST.

Preparation of anhydrous pinacol. HAROLD S. KING and W. W. STEWART. *Proc. Trans. Nova Scotian Inst. Sci.* 17, 262-7 (1930)—The dehydration of pinacol hydrate was carried out by 3 methods: (1) To 118 g. pinacol, 108 g. water, and 700 cc. Et₂O cooled and shaken, 250 g. coarsely powd. CaCl₂ was added. After standing, the mixt. was filtered and the filtrate dist'd with a 3 ft. fractionating column. The fraction b. 170-4° weighed 108 g. and was pure anhyd. pinacol. In all, 90% of the hydrate was recovered. (2) Pure pinacol hydrate dist'd on an oil bath under the same conditions gave only 75-80% of anhyd. pinacol. (3) Pinacol (118 g.), 108 g. water, and 200 cc. CCl₄ were dist'd at 68° until all the water was removed through a condenser connected with an automatic separator from which the CCl₄ siphoned back to the distg. flask. The CCl₄ soln was then heated on a metal bath to 170°. Anhyd. pinacol in 96% yield remained in the flask. The CCl₄ dehydration method was used for the analysis of pinacol hydrate. It was found to contain 4.9% uncombined water when dried in the air for 24 hrs. RACHEL BROWN.

Identification of fatty acids. I. REINHARD SEKA and ROBERT H. MÜLLER. *Monatsh.* 57, 97-105 (1931).—The reaction with *o*-C₆H₄(NH₂)₂ is proposed as a method of identifying fatty acids; the method seems to be applicable to acids up to C₁₁; the possible use with unsat'd acids is being studied. *o*-C₆H₄(NH₂)₂ and PrCO₂H, heated at 140-50°

acids with 18 and 22 C atoms is nearly the same as that of the corresponding *trans* ethylenic acids, while it is smaller than that of the *cis*-ethylenic acids. Moreover, parachors of the satd acids are greater than those of the corresponding unsatd acids while the parachors of the latter are the greater in the *cis*-forms. To complete the investigation the parachor of 2,3-oleic acid (VII) was detd, and since it was practically equal to that of II, the former has the *trans*-configuration as indicated by earlier exp (cf S, C A 19, 1245). The differences in the parachors of each pair of ethylene isomers are small, but since these differences are systematic in both series and for different temps they are not exp'd errors. Surface tensions γ were detd by the method of measuring boiling pressure with the Sugden app (cf C A 16, 2627; 18, 1771), and the parachors were calcd by the formula $P = M\gamma^{1/4} / (D - d)$ of Sugden. The following d give d, γ and the parachor of the acids at 90° and at 95°, resp. I, 0.848, 0.84404, 27.94, 27.52, 765.07, 766.93, II, 0.84370, 0.83670, 26.56, 26.31, 760.44, 764. VII, 0.84843, 0.84039, 27.00, 26.89, 759.31, 763.70, V, 0.83602, 0.83102, 26.89, 26.777.79, 778.87 (cf 778 of Hunsen and Maass, C A 23, 1549), IV, 0.8357, 0.8327, 28.27.77, 932.40, 938.91, III, 0.8324, 0.8294, 27.40, 27.28, 929.20, 934.36, VI, 0.8223, 0.8227.77, 37.61, 950.07, 951.36. The exp'd values of the parachors are not comparable with those calcd theoretically, because the theoretical parachors are always different from the exp'd values when the no. of C atoms comprising the chain is very high. The results show that I and IV have *cis*-configurations, and II, VII and III have the *trans* configurations. C. C. DAVIS

Separation of glycerides. XVI-XVIII. BUNSUKE SUZUKI, *Proc. Imp. Acad. (Japan)* 7, 9-11 (1931), cf. C A. 23, 4928.—XVI. With YUTAKA MASUDA.—By the principle of fractional pptn with CaCl₂ from EtOH and extn of the petr eth insol portion with various solvents the brominated oil from the bones of the common finch (*Balaenoptera physalus*), the following bromides were isolated (only m p give dilinolenol-zoomarin, 65°, clupanodono-linolenol-zoomarin, 132°, clupanodono-arachidonol-olein, 93°; C₂₁H₃₉O-clupanodono-linolenin, 123°, arachidonol-linolenol-olein, 14 triarachidonin, 218°; the following are oils: linolenol-dizoomarin, stearyl-linolenol-zoomarin, 172°; distearyl-linolenol-zoomarin, 188° (decompn.), tristearidonin, 1 (decompn.), triolein and dierythro(?)linolein. XVII. With SHIRO OGI AND UZO NISHIMOTO.—The skin of the finch gave the following bromides: C₂₁H₃₉O-linolenol-steardonin, 122°, (C₂₁H₃₉O)-steardonin, 143°; cetol-linolenol-steardonin, 172°; disteardonol-zoomarin, 188° (decompn.), tristearidonin, 1 (decompn.), triolein and dierythro(?)linolein. XVIII. With RYOZO YONEDA AND UZO NISHIMOTO.—The internal organs of the finch gave the following bromides: clupanodono-linolenol-linolenin, 105°; C₂₁H₃₉O(?)linolein, 93°; clupanodono-steardonol-zoomarin, 193°, diarachidonol-gadolein, 220° (decompn.); linolenol-dizoomarin, or C. J. WES

The structure of synthetic mixed triglycerides. RAMKANTA BHATTACHARYA AND THOMAS PERCY HILDRICH, *Proc. Roy. Soc. (London)* 129A, 468-76 (1930).—Synth triglycerides were prep'd from various combinations of lauric, palmitic, and stearic acids with unsatd acids (85% oleic and 15% linoleic). The curve of the molar percentage of the fully satd glycerides plotted against the molar percentage of satd acid in the synthetic mixt. gave corresponding values for a no. of natural animal fats (including butter and tallow) and for some vegetable pericarp fats, while vegetable fats have somewhat less fully satd glycerides than the synthetic fats for a given ratio of satd to unsatd acids. Vegetable seed fats are almost invariably assembled on different principles, the glyceride structure being detd by a tendency to even distribution of the acids; fully satd glycerides rarely appear in appreciable quantity until the molar proportion of satd acids in the mixed fatty acids reaches 60%. H. J. DEUEL, JR.

Origin of stereochemistry. EMANUELE PATERNO, *Gazz. chim. ital.* 60, 911 (1930).—Certain misstatements in a recent book by Oddo ("Trattato di chimica organica") are discussed. C. C. DAVIS

Stereochemistry of organic compounds. II. Spatial rearrangement of the α in the pentaerythritol molecule. LUDWIG ORTHNER AND GERTRUD FREYSS, *A* 484, 131-54 (1930); cf. C A. 22, 81.—Details are given for the prep'n of the mono- and diacetone (II) derivs of pentaerythritol (III) (cf C A. 22, 1327). I and Na C₆H₅Me, heated 3 hrs., then treated with MeI and again heated 2 hrs., give 43% of mono-Me deriv (IV), b_p 129-30°; Ag₂O and MeI with I give 70% of III; p-nitrobenzoate m. 90°. IV with NaOH in 70% MeOH gives 60% of the Me deriv. of III, m. 72°, in Me₂SO, and 5% KOH give a mixt. of the di-Me deriv. (V) of III, b_p 139°, m. 103° (dibenzoylate, m. 70°), and the tri-Me deriv. (VI) of III, b_p 103-4° (p-nitrobenzoate, m. 53-4°) and BrCl in C₆H₅N give the di-Bz deriv. (VII) of III, m. 110°, hydrolyzed by HCl in Me₂CO the di-Bz deriv. (VIII) of III, m. 75°; further action of BrCl gives the tetra-Bz deriv.

III, m. 94° , identical with that obtained from III and BrCl . I, AgO and AcONa in CHCl_3 , heated 1 hr., give 90% of the *di-Ac deriv.*, m. $49-9^{\circ}$; hydrolysis at room temp. gives 25% of the *di-Ac deriv.* of III, b.p. 159° ; on standing, the tetra-Ac deriv. (VII) and III sep. Cond. measurements are reported for I, V and VI in H_2O and H_2BO_3 . Dipole moments are given for VII and II; μ was found to be $2.19 = 0.03$ and $2.26 = 0.07$, each $\times 10^{-18}$ e. g. s. The spatial structure of III is discussed and illustrated by diagrams. C. J. WEST

Preparation of 1,3,9- and 3,7,9-trimethyluric acid. HEINRICH BILTZ and HEDER PARDON. *Ber.* 63B, 2576-80(1930)—*o*-Methoxytheophylline (I), from the 8-Cl compd. with Na in MeOH, m. 283° , at once resolidifying (probably rearranging into 3,7,9-trimethyluric acid (II)) and then m. $370-80^{\circ}$. Heated 0.5 hr. at 200° (not higher), 10 g. I gives 5 g. II, m. $373-5^{\circ}$ (decompn.), readily methylated by alk. Me_2SO to tetramethyluric acid, m. 225° . 1,3,9-Trimethyluric acid (III) is prepd. by heating K 1,3-dimethylurate with $\mu\text{-MeC}_4\text{H}_9\text{SO}_3\text{Me}$ in $\phi\text{-C}_6\text{H}_5\text{Cl}$ at 180° ; the crude product is purified by removing the more strongly acid 1,3-di-Me compd. by boiling up in dil. NH_4OH or by just dissolving in 2% KOH and pptg. the III with CO_2 ; yield, 3.2 g. from 10 g. of the di-Me acid. The III decompn. 340° , as does that prepd. from 9-methyl μ -uric acid (Strufe gives 347°). An attempt was made to prep. III from 8-chlorotheophylline (obtained in 8 g. yield in needles, m. 304° (decompn.)), from 10 g. theophylline with KClO_4 in HCl , but the Cl compd. was unchanged by KOH or NaOH in boiling MeOH or by KOMe or NaOMe at 140° in sealed tubes. C. A. R.

Preparation and properties of crystallized alkali salts of L-cystine. GERHART TONKES and THEODORE F. LAYTON. *J. Biol. Chem.* 90, 213-13(1931)—The Li, Na and K salts of L-cystine have been prepd. by pptn. of alc. alk. cystine solns. with MeCN. A method of recrystn. was not found and reprecip. under the conditions of prepn. did not improve the purity or cryst. form. However, analysis indicated that the original preps. were fairly pure. The salts are very sol. in water and the aq. solns. are practically stable for several days, a noteworthy fact in view of the known instability of cystine solns. when boiled with excess alkali. [soln. calcd. for cystine at 20° , for the Li, Na and K salts, resp., in water is -94.4° , -90.9° and -91.1° . A. P. LOTTNER

Condensation between aliphatic esters and ketones. S. O. FOWELL and KERR M. SEYMOUR. *J. Am. Chem. Soc.* 53, 1049-51(1931)—With the Claisen condensation, AcOEt and MeCOEt or EtCOEt and MeCO give 11 and 15% resp. of $\text{MeCOCH}_2\text{CH}_2\text{CH}_2\text{COCH}_3$, EtCOEt and MeCOEt or PrCOEt and MeCOEt give 9 and $\frac{1}{2}$ of 3,5-(EtCOCH_2) $_2$, AcOEt and MeCOCH_2Me or iso-PrCOEt and MeCO give 14 and 6% resp. of 2-methyl-3,5-hexanedione. Modifications of proportions and conditions did not change the course of the reaction in the cases studied. C. J. WEST

Sugar oxidation and destruction. X. Formation of methylglyoxal from sugars and from related substances under the influence of hydrogen peroxide. K. BERNHATER and H. TCHERNIK. *Biochem. Z.* 230, 484-92(1931); cf. *C. A.* 24, 4268.—Under the influence of H_2O_2 in the presence of FeSO_4 in acid soln., $\text{MeCOCH}_2\text{CHO}$ is formed from various sugars and especially from γ -Me glucoside. In the absence of FeSO_4 however, acids, particularly HCO_2H , are produced instead. XI. Formation of phenol-like compounds in the decomposition of sugar. K. BERNHATER and J. NARR. *Ibid.* 493-500.—Phenol like substances are formed in very minute amts. and only under special conditions from carbohydrates. KCO_2H when used in making the phenol fusion mixt. leads to the formation of *o*- and particularly of *m*- $\text{HOCH}_2\text{CH}_2\text{CO}_2\text{H}$ from phenol. XII. The formation of higher fatty acids in the destruction of sugar. *Ibid.* 501-4.—There is no evidence found that higher fatty acids result when sugars are acted upon by concentrated alkalis. S. MORGULIS

Glucuronic acid fermentation. II. TEIZO TAKAHASHI and TOSHIMASA ASAI. *Proc. Imp. Acad. (Japan)* 7, 5-8(1931)—See *C. A.* 24, 4319. C. J. WEST

Course of addition of sodium enol alkylmalonic and sodium enol alkylcyanoacetic esters to unsaturated esters. ARTHUR BLUMHART and JOHN ROSS. *J. Am. Chem. Soc.* 53, 1150-72(1931), cf. *C. A.* 25, 82.—The course of the addn. of a deriv. of the Na enol malonic ester type to an α,β unsatd. ester proceeds according to the law of chem. neutralization. That addn. product is therefore formed in which the positive energy of the Na atom is best neutralized, since the max. energy degradation is thus realized. In the addn. of Na enol alkylmalonic esters to fumaric ester, the alkyl group migrates so that the addendum parts are alkyl and $\text{C}(\text{CO}_2\text{Et})\text{C}(\text{ONa})\text{OEt}$; in the addn. of Na enol alkylcyanoacetic esters to crotonic ester, the addendum parts are similarly alkyl and $\text{-C}(\text{CN})\text{C}(\text{ONa})\text{OEt}$. In the addn. to crotonic ester, the yield of addn. product was greater in the case of Na enol methylcyanoacetic ester (50%) than with Na enol ethylcyanoacetic ester (33%). However, the yield of addn. product is detd.

by several complex factors so that it is not possible from these expts. to make any statement regarding the relative ease of migration of the Me and Et groups in these addn. reactions. The max chem neutralization of the Na is realized by the Na remaining attached to the carbonyl O atom in the malonic or cyanoacetic groups. The Na atom, however, becomes better neutralized than it was in the addendum by the spatial negative influence of the acquired CO_2Et group. Esters of the type $\text{ClHX}(\text{CO}_2\text{Et})\text{ClHY} \cdot \text{CZ}(\text{CO}_2\text{Et})_2$ (where $\text{X} = \text{R}$ or H , $\text{Y} = \text{R}$, H or CO_2Et and $\text{Z} = \text{R}$ or H , also R is an alkyl or aryl radical), apparently form enolic Na derivs contg the group $\text{-ClX} \cdot \text{C}(\text{ONa})\text{OEt}$, but these immediately decompose to form ethylenic α, β -esters and Na alkylmalonic esters. The retrogression of the free ester (A) by Na is probably due to the formation of such an enolic Na deriv. It would follow from this relationship that the Na atom involved in the addn reactions specified above cannot migrate during the addn process. However, if the single CO_2Et group of (A) can form an enolate, an amt. of Na corresponding to the relative acidity of the 2 enolic forms of ester (A) ($\text{Z} = \text{H}$), would migrate to the single CO_2Et group according to the partition principle, and spontaneous retrogression would lead to the reformation of a further quantity of the Na enol deriv. Thus, by this process, in agreement with the partition principle, an apparent balanced state would be achieved between the 2 possible stable Na derivs. It has not been found possible to obtain addn to the γ C atom of a β, γ unsatd ethylenic ester or nitrile. Addn occurs in the case of allyl cyanide and styrylacetic ester at the α - and β -C atoms with migration of an α -H atom. $\text{ClH}_2(\text{CO}_2\text{Et})_2$ and $(\text{ClHCO}_2\text{Et})_2$ with a little EtONa give 90% of the Et ester, b_p 187°, of propane $\alpha, \beta, \gamma, \gamma$ -tetracarboxylic acid, m 153°. The ester (22 g) in an Et_2O suspension of anhyd EtONa , methylated with 10 g MeI , gives 21 g of the Et ester, b_p 180°, of butane- $\alpha, \beta, \gamma, \gamma$ -tetracarboxylic acid, m 170°. Pentane- $\beta, \gamma, \delta, \delta$ -tetracarboxylic acid, m 166°, results from the sapon of the ester obtained from the butane deriv and MeI with EtONa at 0°. The propane deriv and EtI with EtONa give the Et deriv, b_p 180°, of pentane- $\alpha, \beta, \gamma, \gamma$ -tetracarboxylic acid, m 177°. Pentane- $\alpha, \alpha, \beta, \gamma$ -tetracarboxylic acid m 170° (decompn). Hexane- $\beta, \beta, \gamma, \delta$ -tetracarboxylic acid, m 170°, is formed from the Et ester of the last-named acid, MeI and EtONa at 0°. Addn of $\text{MeC}(\text{CN}) \cdot \text{C}(\text{ONa})\text{OEt}$ to $\text{MeCH} \cdot \text{CHCO}_2\text{Et}$ gives 50% of the Et ester, b_p 145-8°, of α, β -dimethyl- γ -cyanoglutaric acid, which could not be crystd, the corresponding γ -carboxy deriv m 142° and is not identical with the acid obtained from tiglic acid, β, γ -Dimethyl- γ -cyanoglutaric acid m 152°. $\text{EtC}(\text{CN}) \cdot \text{C}(\text{ONa})\text{OEt}$ and $\text{MeCH} \cdot \text{CHCO}_2\text{Et}$ give 33% of the Et ester, b_p 153°, of γ -cyano- α -ethyl- β -methylglutaric acid, m 147°, this gives with HCl a mixt of α -ethyl- β -methylglutarimides, m 92° and 102°; hydrolysis of the imide, m 92°, gives α -ethyl- β -methylglutaric acid, m 88°. Complete hydrolysis of the above ester gives α -ethyl- β -methyl- γ -carboxylglutaric acid, m 147°. $\text{MeC} \cdot \text{ClEtCO}_2\text{H}$ and $\text{NCCH} \cdot \text{C}(\text{ONa})\text{OEt}$ give 60% of the Et ester, b_p 152°, of α -ethyl- β -methyl- γ -cyanoglutaric acid, m 132°; complete hydrolysis gives the γ -carboxy deriv, m 141°. Ethylation of γ -cyano- β -methylglutaric ester gives the Et ester, b_p 152°, of γ -ethyl- β -methyl- γ -cyanoglutaric acid, m 139°; γ -ethyl- β -methyl- γ -carboxylglutaric acid could not be crystd. $\text{CH}_3 \cdot \text{CHCH}_2\text{CN}$ and $\text{NCCH} \cdot \text{C}(\text{ONa})\text{OEt}$ give 90% of β -methyl- γ -carboxyglutaronitrile, b_p 160°; $\text{MeC}(\text{CN}) \cdot \text{C}(\text{ONa})\text{OEt}$ gives α, β -dimethyl- γ -carboxyglutaronitrile, b_p 152°. $\text{PhCH} \cdot \text{CHCH}_2\text{CO}_2\text{Et}$ and $\text{CNCH}_2 \cdot \text{CO}_2\text{Et}$ with EtONa give Et β -benzyl- γ -cyanoglutarate, b_p 193°; hydrolysis gives β -benzyl- γ -carboxylglutaric acid, m 158°, the Na deriv of the ester and MeI give Et γ -methyl- β -benzyl- γ -cyanoglutarate, b_p 194°, alk hydrolysis gives γ -methyl- β -benzyl- γ -carboxylglutaric acid, m 177° (decompn), giving α -methyl- β -benzylglutaric acid, m 139°. $\text{PhCH} \cdot \text{CHCH}_2\text{CO}_2\text{Et}$ and $\text{MeCH}(\text{CO}_2\text{Et})_2$ with EtONa give the Et ester, b_p 197°, of α -methyl- β -benzyl- γ -carboxylglutaric acid, which exists in 2 forms, m 197° and 118°; heating above their m p gives 2 α -methyl- β -benzylglutaric acids, m 139° (cf. above) and 97°. PhMeClCHCHO and $\text{CH}_2(\text{CO}_2\text{H})_2$ in EtOH give a mixt of the Et esters, b_p 156° and b_p 175-82°, of γ -methyl- γ -phenylbutenoic acid (I), did not crystallize, and γ -methyl- γ -phenyl- α -carboxybutenoic acid, m 151°. I and $\text{NCCl}_2\text{CO}_2\text{Et}$ with EtONa give Et γ -methyl- γ -phenyl- α -cyano- β -acetylbutyrate, b_p 193°; alk. hydrolysis gives γ -methyl- γ -phenyl- α -carboxy- β -acetylbutyric acid, m 162° (decompn), giving γ -methyl- γ -phenyl- β -acetylbutyric acid, m 88°.

α, β, γ -Trimethylglutaric acids. F. E. RAY. *J. Am. Chem. Soc.* 53, 1174-5(1931); cf. C. A. 22, 945—Polemical with Michael and Ross (C. A. 25, 82) regarding the identity of α, β, γ -trimethylglutaric acid.

α, β, γ -Trimethylglutaric acids. ARTHUR MICHAEL AND JOHN ROSS. *J. Am. Chem. Soc.* 53, 1175-6(1931).—Reply to Ray (cf. preceding abstr.).

A supposed asymmetry of meso-tartaric acid. K. SCHERINGA. *Pharm. Weekblad* 68, 143-5(1931)—A theoretical discussion of the spatial configuration of meso-tartaric

C. J. WEST

C. J. WEST

C. J. WEST

C. J. WEST

C. J. WEST

C. J. WEST

C. J. WEST

acid and a possible asymmetry due to a mutual repulsion of the 2 carboxyls.

A. W. DOX

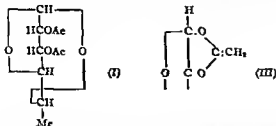
Iron nitrosyls and their behavior on oxidation. HANS REITHLEN, E. ELSEN AND J. EVERET, *Ann* 485, 43-52(1931) — $(\text{NO})_2\text{Fe}(\text{S}_2\text{COEt})_2$ (I) (Manchot and Davidson, *C. A.* 23, 4417) results in 91% yield (on the basis of the Fe) or 95% yield (on the basis of NO) from 10 g FeSO_4 in 70 cc. H_2O and 12-14 g K xanthate in 30 cc. H_2O to which are added 50 cc CS_2 or CHCl_3 and the whole satd. with NO (av. absorption, 1500 \times 50 cc), expts at 0-40°, addn of KOH or AcONa , or variation of the ratio of K xanthate to Fe from 0.5 to 4, always gave the same result. With EtOCOSK there results 35% of dinitroso ferrous Et thiocarbonate (II), $(\text{ON})_2\text{FeSCOS}_2\text{Et}$, brown black, much less stable than I and reacts much more quickly with acids and alkalis. I, on heating at 60-80°, gives a N free sublimate, $\text{S}_2\text{O}_4\text{-di-Et thiocarbonate}$, b_p 76°, n_D^{20} 1.5375. The behavior of I and II on autoxidation is shown in curves. I in C_6H_6 , on shaking with 1 atom of O, gives a compd contg about 30% of $(\text{NO})_2\text{O NOFe}(\text{S}_2\text{COEt})_2$.

C. J. WEST

Carbohydrates. X. Further 1,2-derivatives of glucose. PEACY BRIGL AND RICHARD SCHNITZLE *Ber* 63B, 2884-7(1930), cf. *C. A.* 24, 4509 — An entirely different and materially smoother method of prepg 2 methylglucose (I) (*C. A.* 24, 74) is described. Tetrabenzoylglucose di-Et mercaptal reacts with extraordinary ease with MeI and Ag_2O , yielding half its wt of tetrabenzoyl-2 methylglucose di-Et mercaptal, m 88-9°, $[\alpha]_D^{25}$ 64.33-64.57° (Me_2CO), saponf by NaOEt to 2 methylglucose di-Et mer. captal, m 178°, identical with that obtained from glucose mercaptal with MeI and yielding with aq HgCl_2 I, m 158°, shows slight mutarotation in H_2O , $[\alpha]_D$ changing from 56.6-59.2° in 0.5° hr to 64.9-65.0° in 48 hrs; phenylhydrazone, m 177°, $[\alpha]_D^{25}$ -13.3° ($\text{C}_6\text{H}_5\text{N}$) (heating with more PhNHNH_2 , gives the Me-free glucosazone, m 205°).

C. A. R.

Sugar anhydrides. II. Action of trimethylamine on acetobromo- β -rhamnose. FRITZ MICHAEL AND MARSHA MICHAEL, *Ber* 63B, 2862-6(1930); cf. *C. A.* 23, 3669 — It was shown recently (*C. A.* 24, 2727) that acetobromosugars whose substituents on the 1- and 2 C atoms are in the *cis*-position can form quaternary NH_4 salts with NMe_3 , while the *trans* derivs do not. The acetohalogen derivs of α mannose and β -rhamnose form no quaternary derivs, but while the mannose compd can be recovered almost unchanged even after long continued action, acetobromorhamnose undergoes another reaction. In abs alc after a sufficiently long time almost all the Br is split off and seps as $\text{NMe}_3\cdot\text{HBr}$. Tetraacetyl- β -rhamnoside is not formed, however; there are obtained chiefly a diacetyl- β -rhamnose anhydride (I), 2,3,4-triacetyl β -rhamnose (II) and an unsatd non-cryst compd. The yields of these products vary considerably and it has not yet been possible to establish the conditions necessary for obtaining I, especially, in reproducible yields. I, to which the accompanying structure is assigned provisionally, is probably formed by splitting off of Br with the Ac group on C atom 4 as AcBr which reacts with the alc to form AcOEt and HBr , the latter sepg as $\text{NMe}_3\cdot\text{HBr}$. By using C_6H_6 instead of EtOH as solvent, it is possible to isolate the very unstable $\text{AcBr}\cdot\text{NMe}_3$. I reduces Fehling soln only very little after sapon of the Ac groups but powerfully after boiling with acids, and it is stable toward Br in AcOH and



KMnO_4 in Me_2CO . It has not been possible to obtain the corresponding carbohydrate in pure form, however. Cautious alk sapon gives sirups which after long standing in the air reduce Fehling soln and deposit small quantities of β -rhamnose hydrate, while a dil soln of HCl in MeOH produces a change in rotation from -24.1° to -9.8° but yields no cryst rhamnoside acetate. The simultaneous formation of II renders the isolation of I difficult. The mother liquors from I and II give a sirup which reacts

strongly with Br and KMnO_4 , it may be a compd. of the glucal type or a ketene acetal-like product (III) derived from the tautomeric form of acetobromorhamnose. The easy elimination of AcBr from acetobromorhamnose indicates that the mobile 6-membered pyran ring is fixed in a position in which the Br atom is close in space to an Ac group; if no Walden inversion occurs at C atom 1 in the ring closure, this can result only in the formation of ($\alpha,4$)($\beta,5$)-rhamnosan 2,3-diacetate (I), since the Br is in the *trans*-position to the Ac groups on C atoms 2 and 3. The acetohalogenomannoses, which under the same conditions are stable toward NMe_3 , have an AcOCH_2 group instead of the end Me group of the rhamnose. This end AcOCH_2 group probably favors a form of the pyran ring in which the Br and Ac groups are not close to each other. I (0.9 g from 2 g acetobromo-*l* rhamnose), m 124–5°, $[\alpha]_D^{25} -21^\circ$ (CHCl_3). C. A. R.

Nitrogen-containing sugars. II. Synthesis of peptide-like compounds from amino sugars and amino acids. I. Glucosamine as the component. ALFRED BERTHO, FRITZ HÖLDER, WERNER MEISER and FRANZ HÖTHER. *Ann* 485, 127–51 (1931); cf. C. A. 24, 3761.—Glucosamine-HCl (8.64 g) in 400 cc 0.2 N NaOH, treated during 20 min with 4.52 g ClCH_2COCl with cooling give about 5% of chloroacetyl *N*-glucosamine, m. 163–9°, $[\alpha]_D^{25} 24.8^\circ$ (MeOH), Fehling soln. is easily reduced. The α -bromopropionyl deriv (I) m 200–1°, $[\alpha]_D^{25} 52.5^\circ$ (H_2O) after 17 min, changing to a const. value of 35.2° in 18 hrs, crystn. of this prepn tends to increase the component with the higher rotation, since the const. value for the product crystd. once is 23.78°. The α -bromoisocaproyl deriv (II), m 178° (decompn), also shows mutarotation, the value of $[\alpha]_D^{25}$ changing from 39.67° (H_2O) after 30 min to 24.70° after 24 hrs. I and MeOH-NH₂, on standing 3 months at 0°, give a mixt. of dehydroalanyl *N*-glucosamine anhydride (III), m 272° (decompn), and alanyl *N*-anhydroglucosamine anhydride (IV), m. 199°. Concd. NH_4OH at 0° gives principally III but a small amt. of IV, MeOH-NH₂ at 100° gives almost completely IV. III reduces Fehling soln. only after long heating; it does not liberate N in the Van Slyke app., $[\alpha]_D^{25} -38.0^\circ$ (H_2O), it is unchanged after heating with MeOH-NH₂ 3 hrs at 70–5° and 3 hrs at 100°. IV likewise reduces Fehling soln. only after long heating and in the Van Slyke app. liberates 42% of its N after shaking 7 hrs. The *tetra* Bz deriv of I m 189°, $[\alpha]_D^{25} 86.7^\circ$ (CHCl_3), does not reduce Fehling soln.; a small amt. of *tetrabenzoyl*-(α -hydroxypropionyl)-*N*-glucosamine is also formed, m. 238°, $[\alpha]_D^{25} 15.87^\circ$ (CHCl_3). II also yields a *tetra*-Bz deriv, m. 189°, $[\alpha]_D^{25} 103.1^\circ$ (CHCl_3). Hippuryl *N*-glucosamine, m 200° (decompn), it reduces Fehling soln. but does not react with MeI, $[\alpha]_D^{25} 43.47^\circ$ ($\text{C}_6\text{H}_5\text{N}$). *p*-Bromohippurazide gives the compd. $\text{C}_{20}\text{H}_{24}\text{O}_8\text{N}_2\text{Br}_2$, m. 214°. Benzoylalanine azide gives *benzoylalanyl-N*-glucosamine, m. 222° (decompn), $[\alpha]_D^{25} 59^\circ$ (H_2O). C. J. WAST

Acetylmannoses. VI. The ring structure of the mannose pentaacetates. P. A. LEVENE and R. STUART TIPSON. *J. Biol. Chem.* 90, 89–98 (1931); cf. C. A. 18, 78, 1274; 21, 1969, 3184.—Hudson (C. A. 24, 2726) has recently attributed different ring structures to α - and β -mannose, and if this assumption is correct the α - and β -pentaacetates should also possess different ring structures. L. and T. find on examg the 2 forms that this is not the case. Fifty g of β -mannose pentaacetate treated with AcOH and HBr gave 32 g of 1-bromotetraacetyl-*d*-mannose (I), m. 53–4°, $[\alpha]_D^{25} 123.2^\circ$ (in CHCl_3). Treatment of I (25 g) with Zn dust and AcOH gave 12 g triacetylglucal, m. 54–5°, $[\alpha]_D^{25} -16.3^\circ$ (in abs EtOH). The action of AgOAc on 17 g of I gave 15 g of a compd. m. 74, $[\alpha]_D^{25} +56.6^\circ$ (in CHCl_3), which was found to be identical with α -*d*-mannose pentaacetate. Dry NaOMe acted on I to give the peculiar tetraacetylmethylmannoside of Freudenberg (C. A. 24, 4509), m 104–5°, $[\alpha]_D^{25} -22.6^\circ$ (in CHCl_3), while treatment with Ag_2CO_3 gave 2,3,4,6-tetraacetylmannose (II), m 93°, $[\alpha]_D^{25} +26.3^\circ$ (in CHCl_3). II strongly reduced hot Fehling soln., and on acetylation gave β -mannose pentaacetate, m. 117°, $[\alpha]_D^{25} -24.1^\circ$ (in CHCl_3). It thus appears that the 2 pentaacetates are the α - and β -forms of the same ring isomer. H. W. LEAHY

The preparation and melting point of anhydrous β -maltose. J. GILLIS. *Natuurw. Tijdschr.* 12, 193–9 (1930).— β -Maltose hydrate is dehydrated *in vacuo* below 100°, contrary to the results of deBruyn and van Leent (*Rec trav chim* 13, 218 (1894)). The m-p curves show that the anhyd substance m 108°, while the hydrated form m. 102–3°. The product obtained by deBruyn was probably a supercooled mixt. of about 56% of β - and 46% of α maltose. R. C. HOCKETT

Autohydrolysis of diastatic dextrins. JEAN EFFRONT. *Compt. rend.* 192, 198–201 (1931).—Reducing dextrins are formed during one of the primary phases of the process

of starch saccharification. The maltose encountered in the course of the dextrin disappearance results from the breaking down of these reducing dextrans. In the process of purification by solvents or dialysis autohydrolysis plays a dominant part, while the phenomenon of polymerization also occurs. The final products are, therefore, always impure, a condition similar to the one encountered in the study of the decomposition of starch.

B. S. LEVINE

The supposed depolymerization of glycogen. ENDRE BERNER. *Ber.* 63B, 2760-4 (1930)—R showed recently (*C. A.* 24, 5730) that the supposed depolymerization of inulin described by Vogel and Pietet and by Pringsheim, Reilly and Donovan is explained by the fact that inulin eagerly adsorbs such substances as alc., glycerol, glycol and AcNH_2 , and thereby becomes easily sol. in H_2O , analysis of the supposed depolymerizates showed that the adsorbed substances, together with the pure inulin, corresponded quantitatively to the relatively large observed $f-p$ depressions. B now shows that this is also true of Reilly, Pringsheim and Donovan's glycogen preps. (*C. A.* 24, 3996). When well dried glycogen is dissolved in anhyd. solvents, like fused AcNH_2 , HCONH_2 , or glycol and pptd. with abs. alc., it adsorbs large quantities of these substances, the adsorption being greatest when the H_2O is removed as completely as possible, e. g., if the glycogen soln. is heated some time *in vacuo* before the pptn. The quantity of AcNH_2 adsorbed also decreases with increasing quantities of alc. used for pptn. The ppts. give up their alc. only slowly *in vacuo* over CaCl_2 , but in the air they lose it more rapidly because the hygroscopic substances take up H_2O and the adsorption of alc. is thereby diminished. This explains why the H_2O must be removed before good adsorption will take place. A prepn. obtained by dissolving glycogen in HCONH_2 , at room temp., pptg. and washing with abs. alc. and drying 2 days *in vacuo* over CaCl_2 , gave an apparent mol. wt. of 312 in H_2O but contained 9.16% alc. and 0.25% HCONH_2 , after standing several hrs. in the air the apparent mol. wt. was 610 and a couple of hrs. later 960, the prepn. now contained only 2% alc. A prepn. similarly obtained from a soln. of glycogen in HCONH_2 , which had been heated 1 hr. *in vacuo* at $90-5^\circ$ gave an apparent mol. wt. of 255 increasing, after 2 days in the air, to 317, it contained 4.13% alc. and 0.27% HCONH_2 , and an aq. soln. of the same quantities of glycogen, alc. and HCONH_2 , gave an apparent mol. wt. of 364. A prepn. similarly obtained from glycol gave in H_2O the same opalescence as true glycogen solns. and was colored by I in the same way. The glycogen was prepd. from calves' livers and after careful purification (finally by electro-dialysis) formed a friable powder of the compn. $\text{C}_{11}\text{H}_{19}\text{O}_5$ (ash less than 0.02%), $[\alpha]_D^{20}$ 194.5° in H_2O . The better it was dried, the less did it depress the $f-p$ of H_2O , pointing to the presence of alc., which, like H_2O , is removed with great difficulty, only heating at $100-10^\circ$ *in vacuo* gives a product of the compn. $\text{C}_{11}\text{H}_{19}\text{O}_4$, which may account for the large depressions found by R., P. and D., who dried their glycogen only at 80° .

C. A. R.

Cyclopropene series. I. Diphenylcyclopropanedicarboxylic acid, S. F. DARLINO AND E. W. SPANAGEL. *J. Am. Chem. Soc.* 53, 1117-20 (1931)—Di-*Me* α -bromo- β - γ -diphenyl γ -nitroethylmalonate, m. $117-8^\circ$, with AcOK in MeOH gives di-*Me* 1,2-diphenyl 2-nitro-3,3-cyclopropanedicarboxylate, m. 120° , refluxing with MeONa splits off NaNO_2 , giving the di-*Me* ester, m. $140-2^\circ$, of 1,2-diphenyl-3,3-cyclopropanedicarboxylic acid, m. 190° (evolution of CO_2), oxidation with KMnO_4 gives CH_3Br , Br and the ester give di-*Me* 1,2-dibromo-1,2-diphenyl-3,3-cyclopropanedicarboxylate, m. $104-5^\circ$. Melting the acid gives the lactone, $\text{C}_{11}\text{H}_{13}\text{O}_4$, m. $149-51^\circ$, hydrolyzed to 1,2-diphenyl-2-hydroxy-3-cyclopropanedicarboxylic acid, m. $146-8^\circ$, this is not oxidized by O_2 .

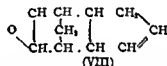
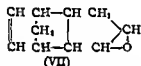
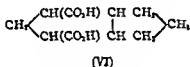
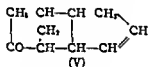
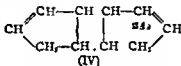
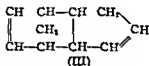
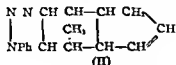
C. J. West

Polymerization and pyrolytic decomposition of phenylacetic anhydride. I. Cyclobutane- α -dione. P. KALININ. *Acta Univ. Latuensis Kim. Fakultat. Seriya I*, No. 13-7, 455-67, 468 (1930) (in German)—In the prepn. and fractionation of $(\text{PhCH}_2\text{CO})_2\text{O}$ there is formed by decompn., a red substance, probably diphenylcyclobutane- α -dione, while $\text{PhCH}_2\text{CO}_2\text{H}$ distils over. This results by the unsym. polymerization of the anhydride with the splitting off of 2 mols. $\text{PhCH}_2\text{CO}_2\text{H}$. The intermediate product of the polymerization is colorless and easily changed into the red diketone. The latter with EtONa undergoes a mol. rearrangement, similar to the benzilic acid rearrangement, the product being $\text{PhCH}_2\text{CHPhCOCO}_2\text{H}$. This is 1 of the few examples of unsym. polymerization, most of the polymerizations yielding sym. compds. The study is being extended to other anhydrides.

C. J. West

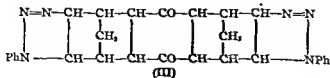
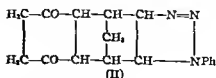
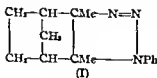
Polymerization of cyclic hydrocarbons. I. Polymeric forms of cyclopentadiene. KURT ALDER AND GERHARD STEIN (WITH HUGO FINZENBACH). *Ann.* 485, 223-46 (1931)—Dicyclopentadiene (I) and PhN_3 give the addn. compd. (II), m. $130-1^\circ$, decompd. by mineral acids with liberation of N_2 . The dihydro deriv. of I, m. $50-1^\circ$, does

not react with PhN_3 , thus indicating structure III as opposed to IV. Oxidation of the dihydro deriv of I with KMnO_4 in Me_2CO gives 3,6-endomethylenetetrahydro-*o*-phthalic acid, m. 137°, and cyclopentane-1,3-dicarboxylic acid, isolated as the anhydride, m. 163°. Decompn. of the bisnitrosocchloride of I gives the unsatd. ketone (V), m. 216°; dihydro deriv., m. 109–10° (semicarbazone, m. 229° (decomps.)); oxidation yields *cis*-pentalan-1,3-dicarboxylic acid (VI), m. 230°. Excess BzO_2H and I give the dioxide, m. 189°, the action of 1 mol. BzO_2H upon I in CHCl_3 at -12° gives a mixt. of *mono*-oxides (VII and VIII), while these cannot be sep'd by crystn., VII adds PhN_3 , giving a *triazole* (IX), m. 149°. From the petr ether mother liquor VIII, b.p. 100–10°, m. 79–80°, is isolated, catalytic reduction of VIII gives a *dihydro deriv.*, m. 118–9°. Tricyclopentadiene adds PhN_3 , giving the compd $\text{C}_{15}\text{H}_{22}\text{N}_4$, m. 199–200° (decomps.).



C. J. WEST

Graduated addition capacity of unsaturated ring systems. KURT ALDER AND GERHARD STEIN, *Ann.*, 485, 211–22 (1931).—The peculiar behavior of dicyclopentadiene toward PhN_3 lead to the study of other compds of this type. Santene and PhN_3 give an *addn compd* (I), m. 86°, Δ^4 -3,6-Endomethylenetetrahydro-*o*-phthalic anhydride gives the compd $\text{C}_{15}\text{H}_{18}\text{O}_4\text{N}_4$, decomps. 225°. Dihydromonocyclopentadienequinone and PhN_3 give the compd II, m. 194°; PhCH_2N_3 gives the compd $\text{C}_{15}\text{H}_{18}\text{O}_4\text{N}_4$, m. 141°, $\text{N}_2\text{CHCO}_2\text{Et}$ gives the compd $\text{C}_{15}\text{H}_{18}\text{O}_4\text{N}_4$, m. 177°. Dihydromonocyclohexadiene- and -butadienequinone do not react with PhN_3 . Biscyclopentadienequinone and PhN_3 give the compd III, darkens above 200° and decomps. 253°, the corresponding hexadiene and butadiene derivs did not react. 3,6-Endoxo- Δ^4 -tetrahydro-*o*-phthalic anhydride and PhN_3 give the compd $\text{C}_{15}\text{H}_{18}\text{O}_4\text{N}_4$, decomps. 200°. *N,N'*-Dicarbethoxy-3,6-endomethylenetetrahydropyridazine and PhN_3 give the compd $\text{C}_{17}\text{H}_{22}\text{O}_6\text{N}_4$, m. 126°.



C. J. WEST

Induction of the reaction between chlorine and benzene and ethylene. T. D. STEWART AND M. H. HANSON *J Am Chem Soc* 53, 1121-8 (1931), cf. C. A. 24, 3984— $C_6H_5Cl_2$, together with substitution products, is formed when C_6H_6 is passed into a C_6H_5Cl soln of Cl in the dark. A mixt of the various stereoisomers is obtained, with the α -form (m 157°) predominating. The molar ratio of total Cl reacting to C_6H_6 reacting varies from 2 to 11 in the present expts. and is greater the lower the partial pressure of C_6H_6 . This is explained on the basis of chain mechanisms for both the C_6H_6 , Cl and C_6H_5Cl , Cl reactions, the latter being induced by and acting as an inhibitor of the former. In homogeneous soln the above ratio, with Cl in large excess, is approx 2. A higher proportion of the Cl is used in substitution than in the heterogeneous reactions. No evidence of a Friedel Crafts type of reaction between the C_6H_5Cl and C_6H_6 , could be found. In a mixt of C_6H_6 , $BzCl$ and Cl no $BzPh$ was formed with introduction of C_6H_6 . C. J. WEST

cis trans-Isomerism and steric hindrance. XII. *o*-Butylcyclohexanols. G. VALLON AND A. GURDON *Bull soc chim* 47, 901-10 (1930); cf. C. A. 24, 1355—A study of the *cis* and *trans*-*o*-butylcyclohexanols (I) and their ethers was undertaken for comparison of their reaction velocities. *o*-Butylcyclohexanone (II) was prepd by catalytic hydrogenation of *o*-Bu C_6H_4 OH (III) followed by oxidation with CrO_3 . The catalytic hydrogenation of III with Pt black yielded butylcyclohexane and 60% of I. The semicarbazone of II m 143-4°, from which was regenerated II, b 90-1°, d_4^{20} 0.914, n_D^{20} 1.4603. Catalytic hydrogenation of the semicarbazone in aq alc. contg HCl yielded the semicarbazide, m 96-7°, oxime of II m 43-4°. II also was obtained by condensation of butanal with cyclohexanone (IV) and subsequent catalytic hydrogenation of the butylidenecyclohexanone in the presence of 5% NaOH according to the process of Grignard and Dubien (C. A. 19, 1217). The fraction b₁₀ 130-45° was not purified but directly dehydrated by distn. in the presence of 2% $(CO_2H)_2$ at 30-40 mm. The fraction now b₁₀ below 100° gave with semicarbazide HCl a compd m 166° which is the semicarbazone of IV. The main fraction b₁₀ 105-20° was directly hydrogenated and yielded II. A 3rd way of obtaining II is by the action of Mg and BuLi on *o*-chlorocyclohexanone. Bu cyclopentyl ketone (V), b₁₀ 104-6°, d_4^{20} 0.893, n_D^{20} 1.4470, is obtained as a by product. The structure of V is proved by its synthesis from cyclopentanealdehyde and BuMgBr and subsequent oxidation of the secondary alc obtained. The formation of a 5-C ring from the 6-C system is emphasized. *cis* I was prepd by Pt black hydrogenation of II in AcOH/HCl. After hydrolysis of the acetate formed it b₁₀ is 101-2°, dimtobenzoate m 76°, acid phthalate, m 58-9°, phenylurethan, m 64-5°. *trans* I, obtained in 87% yield by treating II with Na and alc, b₁₀ 111-2°, d_4^{20} 0.9046, n_D^{20} 1.4654; dimtobenzoate 72-3°, acid phthalate, m 71°, phenylurethan, m 68°. The esterification of the *cis* form is slower than that of the *trans* isomer, hydrolysis with 0.1 N NaOH of the 2 acid phthalates at 39° gives *Ktrans/Kcis* = 3.6 in H_2O , 2.5 in 75% alc. E. G. F.

Oscillating, puckered, centroid model for the benzene ring. MATTHEW L. HUGHES, *J Am Chem Soc* 53, 1182-3 (1931)—Brief discussion. C. J. WEST

Directing influence of certain groups in the benzene nucleus. G. BARGELLINI, *Mem accad Italia (Classe sci fis mat E nat)* 1, Chem. No. 3, 15 pp (1930)—In a preceding paper (C. A. 23, 3910) the influence of OH, O-alkyl, NH, and acylamino groups in directing halogens entering the C_6H_5 nucleus were studied. In the present paper the effect of NH₂ groups when para to each other and blocked by alkyl or acyl groups is studied. Previous work has shown that when an OH group is blocked by an alkyl, and at the *p*-position an NH₂ is blocked by an acyl group, then the first halogen enters a position ortho to the alkoxy, but a 2nd halogen enters ortho to the acylamino group. E. g., $MeOC_6H_4NHAc$ forms successively, acetyl 2-bromo- and then acetyl 2,5-dibromoaniline, m 195°. Its constitution was proved by oxidation which gives 2,5-dibromoquinone. The rule is then that when a deriv of *p*-aminophenol is halogenated, which has an O-alkyl group as well as NHAc, the first halogen enters ortho to the O-alkyl group and the second ortho to the NHAc group. A. W. CONTIERI

Catalytic hydrogenation of organic compounds over copper chromite. HOMER AOKINS AND RALPH CONNOR *J Am Chem Soc* 53, 1091-5 (1931)—Details are given for decomposing $CuNH$ chromite for the prepn of Cu chromite, which is shown to be quite superior in many respects to any catalyst hitherto used. Numerous aldehydes and ketones may be quantitatively and rapidly hydrogenated to the corresponding alcs, this holds true not only for aliphatic CO compds but also for aromatic ones and for keto esters and a keto alc. A carbinol group adjacent to a benzenoid ring may be completely reduced. NO_2 groups may be quantitatively reduced to primary amines without the simultaneous formation of secondary amines. An alkene linkage in a

hydrocarbon as well as in an acid or ketone may be hydrogenated without otherwise modifying the org. compd. Pyridinoid rings and C to N double bonds, as in anils, may be readily hydrogenated. The furanoid ring in furyl alc. may be readily broken with the formation of 1,2- and 1,5-pentanediols in good yields. The new catalyst is not active toward cyanides or toward benzenoid nuclei. It is not nearly so sensitive as Ni to S- or halogen-contg. impurities in the compd. to be hydrogenated. Its lesser sensitivity toward deactivation probably accounts for the fact that it is much more active after it has begun to act than is Ni, despite the fact that Ni is active at lower temps. No special app. such as a reduction furnace is needed and the catalyst need not be freshly prepd. before use. The catalyst ready for use does not change on standing in contact with air or moisture. It is not rapidly deactivated during use. The rate of hydrogenation of Me_2CO over Cu chromite is very much more rapid at higher pressures. With an av. pressure of 35 atms the hydrogenation of 1.73 mols. with 1 g. of catalyst has proceeded to the extent of only 17% in 30 min., while at a pressure of 148 atms 60% of the Me_2CO was reduced and at 212 atms the reaction was 93% complete in 30 min. At the end of 1 hr. the percentage of hydrogenation was 22, 92 and 100 for the 3 pressures given above. Zn chromite has also been used as a catalyst but it is much less active than the Cu compd. and has been found to catalyze condensation. C. J. WEST

Organic syntheses facilitated by pressure. GILBERT T. MORGAN. *Chemistry and Industry* 50, 104-9 (1931).—A review with many references to the patent literature of important syntheses facilitated by pressure was presented by the Herter Memorial Lecturer in which the following were discussed: aromatic amines from phenols and NH_3 , alkylanilines from anilines and alics., alkali fusions of substituted naphthalenesulfonic acids, formation of hydrocarbons, alics., aldehydes, ketones, etc., by pressure reactions in the gas phase from CO and H_2 with various catalysts, dehydration of alics. to form ethers, gaseous polymerization of C_4H_6 , hydrogenation in general, oxidation in general, CO_2 and phenols as in the Kolbe synthesis for salicylates, CO_2 and aromatic hydrocarbons in the presence of AlCl_3 to give benzoic acid, toluic acid, etc., anilines and CO_2 to form CO_2H acids, ketones and urea derivs. In autoclaves pressures from 50 to 300 atms. may be developed without personal risk, providing the temps. are not excessive, above 360° care must be taken. It is best to assume that with every autoclave in const. use there will come a time when failure is to be expected. It is believed that a systematic study of high pressure reactions will in many cases lead to discoveries whereby the use of pressure will be partially or entirely avoided as in the amination of naphthols where a more efficient catalyst has led to a reduction from 50 to 5 atms. Under conditions of high temp. and pressure the oxides of C may be hydrogenated catalytically to yield a complex mixt. of alics., aldehydes, acids and esters, these results are somewhat comparable with the transformations of photosynthesis. The more complex products of vital activity are missing from high-pressure syntheses probably because at the high temp. also involved, these substances are incapable of existence, more efficacious catalysts by causing a lowering in the temp. might make the synthesis of such natural products possible. N. A. LANGE

Transformation of aldehydes into higher molecular amines. A. SKITA AND G. FREIL. *Ann.* 485, 152-73 (1931); cf. *C. A.* 21, 2876.—Previous work has shown that $\text{EtCH}_2\text{N}(\text{C}_6\text{H}_{11})_2$ (I) on reduction with Pt and H gives $\text{PrCHMeCH}_2\text{NHC}_6\text{H}_{11}$ (II). The present study is an extension of this observation. Catalytic reduction of $\text{MeCH}_2\text{N}(\text{C}_6\text{H}_{11})_2$ at room temp. and with a pressure of 3 atm. of H is complete in 0.5 hr., giving 8% of the Et deriv. (*picrate*, m. 231°), and 33.7% of *N*-butylcyclohexylamine, b. $202-4^\circ$ (HCl salt, m. 290°), the latter also results in 51.5% yield by reducing *N*-(2-buten-1-yl)cyclohexylamine, b. $90-1^\circ$. Reduction of I gives 49% of II, II also results from *N*-(2-methyl-2-penten-1-yl)cyclohexylamine, b. $118-9.5^\circ$, in 49% yields, using an equiv. amt. of EtCHO with I increases the yield of II to 76.4%, II HCl salt, m. 139° , phenylurea, m. 99° . Butylcyclohexylamine, b. $85-7^\circ$, is reduced in 5 hrs., giving 75% of *N*-(2-ethylhexyl)cyclohexylamine, b. $140-4^\circ$, this also results in 56% yield by reducing *N*-2-ethyl-2-hexen-1-ylcyclohexylamine, b. $139-43^\circ$; HCl salt, m. 95° , picrolonate, m. $225-6^\circ$ (decompn.). Reduction of isovalercyclohexylamine gives 50% of the *N*-iso-Am deriv., b. $89-93^\circ$ (HCl salt, m. $279-80^\circ$) and 15% of *N*-(2-isopropyl-5-methylhexyl)cyclohexylamine, b. $143-5^\circ$ (HCl salt, m. 105° , acid oxalate, m. $193.5-4.5^\circ$); *N*-(2-isopropyl-5-methyl-2-hexen-1-yl)cyclohexylamine, b. $146-52^\circ$, gives 61.5% of the same amine. Isovalercyclohexylamine is reduced in 3 hrs., giving 20% of the iso-Am deriv., and 16.4% of *N*-(2-isopropyl-5-methylhexyl)ethylamine, b. $213-20^\circ$ (picrolonate, m. $221-2^\circ$ (decompn.)). *N*-Heptylcyclohexylamine, b. $132-5^\circ$ (HCl salt, m. $211-2^\circ$); *N*-(2-amylnonyl)cyclohexylamine, b. $208-10^\circ$ (42.3% yield), HCl salt, m. 90° , acid oxalate, m. 153° , *N*-(2-amyl-2-nonen-1-yl)cyclohexylamine, b. $208-11^\circ$, reduction gives 25.8% of the

amine Reduction of hydrocinnamylcyclohexylamine gives 15% of the 3-phenylpropyl deriv. and about 16% of *N*-(2 benzyl 5-phenylamyl)cyclohexylamine, whose *HCl* salt *m* 169-70° and acid oxalate *m* 172°, this also results in 67% yield by reducing *N*-(2-benzyl 5-phenyl 2-penten-1-yl)cyclohexylamine, *b*₁₁ 265-70° (partial decomposition) Catalytic reduction of a mixt of I and AcI in AcOH gives, in addn to $\text{C}_{11}\text{H}_{17}\text{NH}_2$, its Ft. Bu and 2-methylamyl derivs, 29.5% of *N*-(2 methylbutyl)cyclohexylamine, *b*₁₁ 90-101° (*HCl* salt, *m* 234°, phenylurea, *m* 118°), this also results in 61% yield by reducing *N*-(2 methyl-2 buten-1-yl)cyclohexylamine, *b*₁₁ 100-5° I and BrII give 11% of *N*-(2-methyl 3-phenylpropyl)cyclohexylamine, *b*₁₁ 165-72° (*HCl* salt, *m* 184°, phenylurea, *m* 116°), this also results in 60.5% yield from *N*-(2 methyl 3-phenyl 2-propen-1-yl)cyclohexylamine, *b*₁₁ 168-70° I and lufural give 8.7% of *N*-(2 methyl-3 furylpropenyl)cyclohexylamine, *b*₁₁ 167-75° (*HCl* salt, *m* 189°); 0.7% of its dihydro deriv, *b*₁₁ 156-61°, and 8% of the tetrahydro deriv, *b*₁₁ 143-50° (acid oxalate, *m* 165°) The same compds are obtained on reducing *N*-(2-methyl-3 furyl 2-propen-1-yl)cyclohexylamine, *b*₁₁ 158-62°, in 7.3, 1.8 and 10.75% yield, resp. Furfuralcyclohexylamine, *b*₁₁ 131.5-5.5°, reduction gives the *N* furanmethyl deriv, *b*₁₁ 123-4°; phenylurea, *m* 113°. Isobutylal-cyclohexylamine and *HClHO* give 61.8% of *N*-(2,2-dimethyl-3 hydroxypropyl)cyclohexylamine, *b*₁₁ 123-5°, *m* 38°, *HCl* salt, *m* 221°, phenylurea, *m* 150°. di-Ac deriv, *b*₁₁ 198-0° This also results in 37.8% yield by reducing *N*-(2,2-dimethyl 3 propanol 1-yl)cyclohexylamine, *b*₁₁ 103-10° These results indicate that the reaction proceeds through the hydrolysis of the amine to the aldehyde, which then forms a Schiff's base with the amine and as such is reduced Only in the case of I is practically complete hydrolysis observed The reaction does not occur with aromatic aldehydes of the type of BrII but does with the fatty-aromatic aldehydes, such as $\text{PhCH}_2\text{CH}_2\text{CO}_2\text{H}$

C. J. WEST

The manufacture of acetaldehyde. A. ULLMANN. *Metallber* 21, 315-6, 363-4, 413(1931)—A review of patents

E. H.

Removal of hydrogen halides from organic halides. C. R. NOLLER AND R. DIMMICK. *J. Am. Chem. Soc.* 53, 1185-6(1931)—N. and D. report that they have been investigating the reaction of org. halides with tert. amines (cf Semb and McElvain, *C. A.* 25, 1252, who report the action on sec. amines) The main reaction other than quaternary salt formation appears to be the removal of halogen acid, although rearrangement of the salt to alkylpyridine hydrohalides has not been excluded

C. J. WEST

Fluosilicates of organic bases. II. C. A. JACOBSON. *J. Am. Chem. Soc.* 53, 1011 5(1931), cf *C. A.* 23, 117—The fluosilicates of the following compds were prepd and the crystals photographed: *o*-Toluidine, *m* 268-9°, di-*o*-naphthylamine, *m* 218° (decomps) di-*β*-naphthylamine, *m* 236.3°, di-*m*-nitroaniline, *m* 200°, diphenylamine, *m* 169°, *m*-phenylenediamine, *m* 243-4°; *p*-phenylenediamine, decomps on heating; diethylaniline, *m* 165.3°, dinitroazodiphenylamine, *m* 124.5° (decomps); di-*p*-aminobenzene, *m* 220° (decomps), di-*p*-aminobenzoic acid, *m* 240-5°. The soly in EtOH at 25° is given

C. J. WEST

Mechanism of chemical reactions. II. Mechanism of the synthesis of secondary and tertiary amines by reduction. KARL KEDLER, W. FESCHKE AND W. DEHN. *Ann.* 485, 113-26(1931), cf *C. A.* 17, 2278— PhCH_2NH_2 gives almost quant. (PhCH_2)₂NH on reduction with Pd and H₂ in 90% EtOH for 3 hrs. $\text{PhCH}_2\text{CH}_2\text{NH}_2$ (I), similarly treated, gives only 10% of the sec. amine. This reaction is assumed to proceed as follows: $\text{RCH}_2\text{NH}_2 \rightarrow \text{RCH}_2\text{NH} \rightarrow \text{RCH}_2\text{NCH}_2\text{R} \rightarrow (\text{RCH}_2)_2\text{NH}$ A mixt. of I and PhCH_2OH in $\text{C}_6\text{H}_5\text{Me}$, reduced as above for 3 hrs., gives a mixt. of 6% (PhCH_2CH_2)₂NH but principally $\text{PhCH}_2\text{CH}_2\text{NCH}_2\text{Ph}$ (II) Here it is assumed that the PhCH_2OH is reduced to BrII, which then reacts to give $\text{PhCH}_2\text{NCH}_2\text{CH}_2\text{Ph}$ (III) and this is reduced to II. III, prepd from BrII and I, is nearly quant. reduced to II in 5 min. PhCSNH_2 (IV), on reduction with Al Hg in moist Et₂O, gives 85% of PhCH_2NH_2 and 8% (PhCH_2)₂NH, while $\text{PhCH}_2\text{CSNH}_2$ (V) gives 26% of I and 66% (PhCH_2CH_2)₂NH In 55, 90 and 145 min., resp., 84, 94 and 99% of IV and 35, 41 and 31% of V, resp., were reduced at 20° in moist Et₂O Both primary and sec. amines result on reduction of $\text{PhCH}_2\text{CH}(\text{OH})\text{NH}_2$ and $\text{PhCH}_2\text{CH}_2\text{NH}$ Reduction of $\text{PhCH}_2\text{CH}(\text{OH})\text{NMe}_2$ with Al Hg gives $\text{PhCH}_2\text{CH}_2\text{NMe}_2$, *b* 204-8°. In the reduction of PhCH_2CN in AcOH-AcONH₄ with Pd and H₂, the Schiff base is an intermediate product and the final product contains about 60% of (PhCH_2CH_2)₂NH. The amt. of sec. amine formed is much larger in a neutral or alk. medium than in an acid medium. PhCN in AcOH is catalytically reduced at 35° and 2.5 atm. H₂ in 20 min., giving 99% of primary amine (VI), in 96% EtOH, 70% is reduced after 180 min., giving 90% VI and 4% sec. amine. PhCH_2CN is 95% reduced in AcOH after 420 min., giving 50% primary and 40% sec.

amine; in EtOH, 50% is reduced in 180 min., giving 17% primary and 68% of sec amine. PhCH_2CN and Pr_3NH , catalytically reduced in EtOH, give a mixt. of 21% of $(\text{PhCH}_2\text{CH}_2)_2\text{NH}$ and 60% of $\text{Pr}_3\text{NCH}_2\text{CH}_2\text{Ph}$. C. J. West

Derivatives of anesthetics. L. CONIGLIO. *Rend. accad. sci. Napoli* [3], 36, 50-60 (1930).—New derivs were prep'd by condensing anesthetic, $p\text{-H}_3\text{NC}_6\text{H}_4\text{CO}_2\text{Et}$ (I), with $\text{Me}_3\text{CHCHBrCOCl}$ (II), BzCl (III), phthalyl chloride (IV) and AcCl (V). The aim was to study the pharmacol action of the derivs in relation to the new radicals introduced into the mol. *N- α -Bromoisovaleryl deriv.* was prep'd from 0.60 g I and 7.90 g II; mixed in a separatory funnel with about 50 cc of a 2% soln of equal parts NaOH and KOH , the caseous white mass was filtered, washed repeatedly with H_2O and dried. The product was sol in all org solvents, a pure material, m 115° , was obtained by dissolving in acetone and pptg with petroleum ether. *N-Bz deriv* from 6.60 g I and 5.62 g III, according to the Schotten-Baumann method, washed with H_2O and dried in *vacuo* over H_2SO_4 , it crystallizes in white needles from alc, sol in acetone, CHCl_3 , AcOH and Et_2O , insol in H_2O , m 15° . *N-Phthalyl deriv*, $\text{C}_{11}\text{H}_9(\text{CO})_2\text{NC}_6\text{H}_4\text{CO}_2\text{Et}$, from 6.60 g I and 7.12 g IV in benzene and cooled, the white mass filtered, washed with H_2O and dried in *vacuo* over H_2SO_4 , long needles from alc crystallize out, sol in warm AcOH and acetone, very sol in cold CHCl_3 , insol in Et_2O and in H_2O , m 152° . *N-Ac deriv* from 6.60 g I in 20 cc benzene and 3.12 g V warmed on the water bath and cooled, filtered, washed repeatedly with Et_2O and dried over CaCl_2 , sol in alc from which it is pptd, with Et_2O , sol in H_2O , acetone, AcOH , slightly sol in CHCl_3 , insol in Et_2O and MeOH , m 181° . The pharmacol properties are being investigated. P M

***p*-Bromoanilides of isobutyric and isovaleric acids.** MARQUERITE KUHN AND S. M. McELVARY. *J. Am. Chem. Soc.* 53, 1173-4 (1931).—The derivs were prep'd from the acid chloride and $p\text{-BrC}_6\text{H}_4\text{NH}_2$, *p*-bromoisobutyranilide m 150.1° *p*-bromoisovalerianilide, m. 125.0° . C J West

The organic sulfonic and sulfinic acids. JULIUS V. BRAUN AND KARL WEISSBACH. *Ber.* 63B, 2336-47 (1930).—In contrast to the large amount of work which has been done on the imidocarboxyl chlorides, no attempts had ever been made to prep. the imidosulfonyl chlorides, $\text{RSCl NR}'_2$, and the only expt. in the field of the SO_2H acids recorded in the literature indicated that the relations are materially different from those in the CO_2H acid series, PCl_5 reacts with PhSO_2NHPh only after long heating and the $-\text{SO}_2\text{NH}-$ group is not attacked, only chlorination of the Ph nucleus taking place. $\text{PhSO}_2\text{NHIEt}$ also reacts only at 100° and very slowly, forming (probably as the result of chlorination outside of the $-\text{SO}_2\text{NH}-$ group) a very difficultly separable oily mixt boiling in a high vacuum within wide limits ($100\text{--}70^\circ$). The hope that in the absence of aromatic residues the results would be more clear-cut was completely fulfilled. Expts. on monoalkylamides of BuSO_2H , iso- AmSO_2H and cyclohexylsulfonic acid gave the following results: PCl_5 reacts with compds. like $\text{BuSO}_2\text{NHIEt}$ considerably more sluggishly than with the amides of non-aromatic CO_2H acids but, without recourse to extreme conditions, gives compds having the expected compn BuS(O)Cl NEt and whose properties were, at first, surprising, the compds are stable to heat and can therefore be distd; their Cl is strikingly unreactive; with H_2O and alc. they react only very slowly, with amines almost not at all. The action of an excess of PCl_5 results in a very slow chlorination of the C atom adjacent to the $-\text{S(O)Cl}<$ group, but with still more Cl under more vigorous conditions the grouping $>\text{C S(O)Cl N-}$ is finally converted into $>\text{C SCl}_2\text{N-}$, in which the 2 new Cl atoms (but only these) are as reactive as the Cl in the imidocarboxyl chlorides, even cold H_2O converts the compds back into the $>\text{C S(O)Cl N-}$ derivs. It is believed that these properties are due to the "imidosulfonyl chlorides" really having a 3-membered ring structure (I) or, from the standpoint of the modern theory of semi-polar union, the structure II. It had been hoped that this might be confirmed by a detn of the parachor, but the usual parachor constds do not hold for this type of compds and the values obtained differed widely from those calcd. If the peculiar behavior of the "imidosulfonyl chlorides" is due to the presence of O, the SO_2H derivs should show no such anomaly, and as a matter of fact compds. like BuSONHIEt react readily with PCl_5 to give HCl , POCl_3 and liquid chlorinated products reacting violently with H_2O and which cannot be distd without decompn. even in a high vacuum. The prepn. of a no of amides of non-aromatic sulfonic acids was used to study a question in this field which hitherto had been tested experimentally only with PhSO_2H , viz., the dismutation of these acids ($2\text{RSO}_2\text{H} + \text{RSO}_2\text{H} = \text{RSO}_2\text{SR} + \text{RSO}_2\text{H} + \text{H}_2\text{O}$). The distn. without decompn. of the BuSO_2H , iso- AmSO_2H and cyclohexylsulfonic acid made it possible to follow the reaction quantitatively. While the mol of RSO_2H which is oxidized at the expense of the other 2 mols cannot be replaced by *o*- or *p*- $\text{C}_6\text{H}_4(\text{OH})_2$ as O acceptor, it can be replaced by mercap-

tans $2\text{RSO}_2\text{H} + \text{HSR}' = \text{RSO}_2\text{SR} + 2\text{H}_2\text{O} + \text{R}'_2\text{S}_2$. Sulfinic esters and sulfonyl chlorides, unlike the free acids, are relatively very stable but the conversion of the chlorides into thiosulfonic acids can be readily effected with Na mercaptides or, even more conveniently, with dithiocarbamates as O acceptors $2\text{R}_2\text{NCS}_2\text{H} + \text{NHR}' + 2\text{CISOR} = [\text{R}_2\text{NC}(\text{S})\text{S}]_2 + 2\text{CHH NHR}' + \text{R}'_2\text{SO}_2\text{SR}'$. Sulfonyl chlorides react according to the equation $2\text{R}_2\text{NCS}_2\text{H} + \text{NHR}' + 2\text{CISR}' = [\text{R}_2\text{NC}(\text{S})\text{S}]_2 + 2\text{CHH NHR}' + \text{R}'_2\text{S}_2$. The AlkSO_2H derivs were prepd by treating the alkyl bromides in Et_2O with Mg and SO_2 and converting the resulting acids into the chlorides with SOCl_2 , the AlkSO_2H derivs by oxidizing the crude $(\text{AlkSO}_2)_2\text{Mg}$ with KMnO_4 and converting the AlkSO_2H into the chlorides with PCl_5 . The liquid BuSO_2H no longer dissolves clear in H_2O after standing a short time at room temp; at 100° (in N) the disproportionation is complete in about 1 hr and the H_2O insol part $b_{\text{m.p.}} 126-8^\circ$ and has the compn of *Bu butylthiosulfonate*, while the strongly acid, H_2O sol part is *butylsulfinic acid*, $b_{\text{m.p.}} 145^\circ$. *Butylsulfonyl chloride*, light yellow, $b_{\text{m.p.}} 38^\circ$, $b_{\text{b.p.}} 78^\circ$, fumes in moist air, is energetically hydrolyzed by H_2O , is stable in the absence of O, even on heating, but in dry air slowly changes into BuSO_2Cl . The *Et* ester, $b_{\text{m.p.}} 85^\circ$, and *ethylamide*, $b_{\text{m.p.}} 109-10^\circ$, show the same oxidizability and non-dimutation. *iso Am isoamylthiosulfonate*, $b_{\text{m.p.}} 176-80^\circ$. *Isoamylsulfinic acid*, $b_{\text{m.p.}} 176-8^\circ$. *Isoamylsulfonyl chloride*, $b_{\text{m.p.}} 60-2^\circ$, $b_{\text{b.p.}} 91-2^\circ$. *iso-Am ester*, $b_{\text{m.p.}} 98^\circ$, remains unchanged at 100° in N. *Ethylamide*, $b_{\text{m.p.}} 120^\circ$ in a high vacuum. *Cyclohexylsulfinic acid*, yellowish oil solidifying on strong cooling to crystals, m $33-5^\circ$, and having approx the compn of a monohydrate. *Cyclohexyl cyclohexylthiosulfonate*, faintly yellow, $b_{\text{m.p.}} 184-6^\circ$. *Cyclohexylsulfinic acid*, $b_{\text{m.p.}} 178-80^\circ$, readily absorbs H_2O to form a monohydrate, m 92° , the m p given by Borsche and Lange for the acid. *Butylsulfonylthylamide*, $b_{\text{m.p.}} 120-2^\circ$. *Ethylamidobutylsulfonyl chloride*, $b_{\text{m.p.}} 124-6^\circ$. *Isoamylsulfonylthylamide*, faintly yellow oil, *imido chloride*, $b_{\text{m.p.}} 129-30^\circ$ (with 2 mols PCl_5 is obtained a *di Cl compd.*, $b_{\text{m.p.}} 129-30^\circ$, which is as resistant to hydrolysis as the mono-Cl compd and probably has the structure $\text{C}_6\text{H}_5\text{CH}_2\text{N}(\text{C}_6\text{H}_5)_2\text{O}(\text{N})$). *Cyclo-*

hexylsulfonylthylamide, $b_{\text{m.p.}} 183-5^\circ$, m 72° , *imido chloride*, $b_{\text{m.p.}} 131-2^\circ$, m $73-4^\circ$. With 4 mols PCl_5 is obtained a brownish oily compd $\text{C}_{11}\text{H}_{19}\text{NCl}_4\text{S}$, $b_{\text{m.p.}} 140-50^\circ$, quickly converted by cold H_2O into a compd $\text{C}_{11}\text{H}_{19}\text{ONCl}_2\text{S}$, m 150° , in which both Cl atoms are very firmly held. *Cyclohexylsulfonylpropylamide*, m 78° , reacts so slowly with PCl_5 that even after 2 5 hrs at 100° with 2 mols PCl_5 it gives only 60% of a mono-Cl compd $b_{\text{m.p.}} 145-60^\circ$, the *heptylamide*, $\text{C}_{11}\text{H}_{21}\text{SO}_2\text{NHC}_6\text{H}_{13}$, m 72° , behaves in the same way. $p\text{-O}_2\text{NC}_6\text{H}_4\text{H}_2\text{SCl}$, $b_{\text{m.p.}} 125^\circ$, with piperidine piperidylthiocarbamate in Et_2O yields 1 mol. piperidine HCl and a mixt. of $(p\text{-O}_2\text{NC}_6\text{H}_4)_2\text{S}_2$ and piperidylthiuram disulfide which could not be sep'd anywhere nearly quantitatively but distn in a high vacuum gave a product sep'd by crystn. from alc into the $(\text{O}_2\text{NC}_6\text{H}_4)_2\text{S}_2$ and dipiperidylthiourea.



(I)



(II)

C. A. R.

Separation and synthetic preparation of phenol. TADASHI MIZOSHITA. *Abstracts from Kept Central Lab S Manchuria Railway Co* 1929, 36.—High temp coal tar yields only 0.1897% PhOH and 0.569% cresols. 1 PhCl + $2\text{Na}_2\text{CO}_3$ gives 84% PhOH by heating under pressure at 320° for 2 hrs with Cu_2S as the catalyst. V. F. JARINGTON

N-o-Chlorobenzoyl-o-chlorobenzenesulfonamide. E. WESTHEIM. *J. Am Chem Soc* 53, 1172-3 (1931). $\rightarrow \text{ClC}_6\text{H}_4\text{SO}_2\text{NH}_2$ and $\text{ClC}_6\text{H}_4\text{COCl}$, heated at $180-90^\circ$ for 1 hr, give *N*-o-chlorobenzoyl-o-chlorobenzenesulfonamide, m $154-5^\circ$. Attempts to prep from this diphenic sulfide failed, Cu and NaI in AmOH gave an impure compd, m $255-8^\circ$, with 8.67% S, which had a bitter taste. C. J. WEST

Methylation of alcoholic hydroxyl from the standpoint of the electron theory. JULIUS V. BRAUN, ERNST ANTON AND KARL WEISSBACH. *Ber.* 63B, 2847-61 (1930).—The old valence doctrine does not serve to explain why the proximity of a Ph nucleus or a double bond to the HO group in such compds as PhCH_2OH , $\text{PhCH}(\text{CH}_3)\text{OH}$, $\text{CH}_2\text{CH}(\text{OH})$, dioxindoles , *morphine*, *codeine* and the 2-methylmorphine makes the HO group so easy to methylate, as compared with the alcs, with Me_2SO , and alkali, but the electron theory offers a plausible explanation. The pair of electrons in a homeopolar combination XY is shifted nearer to the one or the other of the partners in the union, depending on the chem nature of X and Y and on the near and, to some extent also, the more distant surroundings. The Ph nucleus and the double bond, which

exert an attraction on electrons, must produce in a compd $\text{Ph} \begin{array}{c} \diagup \\ \text{C} \\ \diagdown \end{array} \begin{array}{c} \diagup \\ \text{C} \\ \diagdown \end{array} \text{O II}$ or $-\text{CH} \text{CH} \begin{array}{c} \diagup \\ \text{C} \\ \diagdown \end{array} \begin{array}{c} \diagup \\ \text{C} \\ \diagdown \end{array} \text{O II}$ a shifting of the electrons toward the left, imparting to the end II atom a weak positive charge and making it more reactive. If this basic assumption is correct, it follows that (1) the inductive or attractive influence of a Ph nucleus or a double bond must diminish greatly with the distance from the HO group, (2) an influence as strong as, or stronger than, that of a Ph nucleus or C C bond must be exerted by every dipole with its + side turned toward the HO group

(e.g., the groups $\text{O}=\text{C}^+ \cdot$, $\text{N}^+ \cdot$, etc.). (3) the effect must be most pronounced when in the chain contg the HO group there is an element having the character of a positively

charged ion (e.g., $\text{Ac}^+ \text{N}^+ \text{C} \begin{array}{c} \diagup \\ \text{C} \\ \diagdown \end{array} \text{O II}$). The behavior of the morphine alkaloids mentioned above suggested that it is not the aromatic nuclei or the double bonds which activate the II of the HO group but that under the influence of the Me_2SO_4 the N is converted into a quaternary form and it is the NH_4^+ ion which acts on the HO group. To test these views a long series of comparative expts was made, the values given below represent the % yield of Me ether obtained from the different HO compds under identical conditions: PhCH_2OH (I), 64, $\text{PhCH}_2\text{CH}_2\text{OH}$ (II), 19, $\text{Ph}(\text{CH}_2)_3\text{OH}$ (III), 5, $\text{Ph}(\text{CH}_2)_4\text{OH}$ (IV), 0, PhCM_2OH (V), 0, $\text{PhCH} \text{CHCH}_2\text{OH}$ (VI), 53, 0 $\text{CH} \text{CH} \text{CH} \text{CH} \text{C}$

CH_2OH (VII), 70, $\text{CH}_2 \text{CHCH}_2\text{OH}$ (VIII), >20, PrOH (IX), 0, AmOH (X), 0, $\text{C}_6\text{H}_5\text{OH}$ (XI), 0, $\text{C}_6\text{H}_5\text{N}(\text{CH}_2)_2\text{OH}$ (XII), 100, $\text{Me}_2\text{N}(\text{CH}_2)_2\text{OH}$ (XIII), 100, $\text{PhNMe}(\text{CH}_2)_2\text{OH}$ (XIV), almost 100, $\text{C}_6\text{H}_5\text{N}(\text{CH}_2)_3\text{OH}$ (XV), about 60, $\text{C}_6\text{H}_5\text{N}(\text{CH}_2)_4\text{OH}$ (XVI), 0, $\text{Me}_2\text{N}(\text{CH}_2)_3\text{OH}$ (XVII), 0, 2 dimethylamino-3-hydroxytetralin (XVIII), 100, 2 β -hydroxyethylpiperidine (XIX), 100, dihydrocodone (XX), almost 100, 5,2-Me(Me_2N) $\text{C}_6\text{H}_4\text{CH}_2\text{OH}$ (XXI), 60, tropine (XXII), 0, $\text{Me}_2\text{NCMe}_2\text{CH}(\text{OH})\text{Me}$ (XXIII), 0, $\text{PhSO}_2\text{NMe}(\text{CH}_2)_2\text{OH}$ (XXIV), 0, $\text{NC}(\text{CH}_2)_2\text{OH}$ (XXV), >50, $\text{EtSO}_2\text{CH}_2\text{CH}_2\text{OH}$ (XXVI), 100, $\text{Me}_2\text{SiCH}_2\text{CH}_2\text{OH}$ (XXVII), 50, $\text{Me}_2\text{SiCH}_2\text{CH}_2\text{CH}_2\text{OH}$ (XXVIII), ?, $\text{MeCH}(\text{OH})\text{CO}_2\text{H}$ (XXIX), 15, $\text{PhCH}(\text{OH})\text{CO}_2\text{H}$ (XXX), 70, $\text{PhCH}(\text{CH}_2\text{OH})\text{CO}_2\text{H}$ (XXXI), 40, phthalide (XXXII), 40, phenylparaconic acid (XXXIII), 40. Comparison of compds I-IV shows how rapidly the influence of the Ph on the HO group decreases with the distance between them. This is also true of the 2 unsatd alcs tested (VI, VIII). Satd aliphatic alcs are resistant toward Me_2SO_4 and alkali (IX, X and XI). The furan ring is similar to the C_6H_5 ring (VII). CN and the more strongly dipolar SO_2 have as great or a greater effect than the aromatic nucleus of the C C bond (XXV, XXVI). In compds with basic N together with HO, the effect is greater than in aliphatic aromatic alcs (XII compared with II, XIII and XIV with III). The influence of N is still strong on a δ -HO (XV) but is no longer shown on ϵ -HO groups (XVI, XVII). That the influence of the N is due to its forming a quaternary salt is indicated by the behavior of XXI and XXIV; its greater influence, as compared with that of the Ph nucleus or C C bond, is shown by the fact that XX is methylated to about the same extent as codeine. The non-methylation of XXII is thought to be due to the branching of the chain between the N and the HO, the base XXIII and the carbinol V behave similarly. That it is not the secondary nature of the HO or the position of the N in the ring that prevents the methylation of XXII is shown by the behavior of the alkamines XVIII and XIX. The electron attraction

should be weakened by increasing the vol of the ionized atom $\text{X}^+ \begin{array}{c} \diagup \\ \text{C} \\ \diagdown \end{array} \text{O II}$, to test this point, the sulfonium compds XXVII and XXVIII were studied but no clear picture of the methylation of XXVIII could be obtained, XXVII, however, acted as had been expected. Of the simple aliphatic HO acids, only XXIX gave results which were at all definite, they indicated that the influence of the C O is not inordinately great, as also indicated by comparison of XXX with I and of XXXI with II. In the reaction between Me_2SO_4 and an activated HO group it may be assumed that as the result of an attraction of the positively charged S by the weakly negatively charged O in the OH group

there is formed an intermediate addn product $\begin{array}{c} \text{R} \\ \diagup \\ \text{H} \end{array} \text{O} \text{SO}_2(\text{OMe})_2$ which then in

some way undergoes further reaction. The extraordinary ease with which alkamines and HO acids are methylated is of practical interest for purposes of prepn and in questions of structure. The HO bases prepd from compds of the type $\text{ArCH} \text{CH}-$ by addn of Br, boiling with H_2O and treatment with amines are formed through an oxide

ArCH CII O which may give either ArCH(NHR)CH(OH)—or ArCH(OH)CH(NHR)—

No means of distinguishing between these two possibilities was hitherto available, but methylation will give an ether which must be identical with or different from the compd ArCH(OMe)CH(NMe₂)—obtained by treating ArCHBrCHBr— with MeOH, then with NH₂Me and methylating exhaustively. By this means it was found that the hydroxymethylamino base (XXXIV) obtained by Mannich from isosafrole really has the structure C₁₁H₁₂O₂N₂CH(NHMe)CH(OH)Me which he assigned to it without any real proof. The methylations were carried out under as nearly identical conditions as possible, the substance dissolved or suspended in 10 parts H₂O at 5° was treated with 6 mols Me₂SO, and 14 mols 10% NaOH, shaken 5 min and allowed to stand (the temp rose to 50–55°) 15 min. The methiodide of XII forms a red chloroplatinate, m. 228°, and a chloroaurate, m. 247°, the methiodide of the Me ether gives an orange chloroplatinate, m. 234° (decompn), and a chloroaurate, m. 109°. Methiodide of the Me ether of XIII, m. 221°, chloroplatinate, decompn 235°. Methiodide of XIV, m. 132°. Me ether, m. 110–21°. Methiodide of XV Me ether, semi-ol., XVIII, m. 31°. Me ether methiodide, m. 223°, methochloride, m. 230°, chloroplatinate, light orange, m. 229°. XXIX with Me₂SO, forms a quaternary salt which cannot be pptd out, the compn of the chloroplatinate, C₁₁H₁₄O₂N₂Cl₂Pt, m. 212°, shows that both the HO and NH groups are methylated, with MeI is obtained a methiodide, decompn 235°, the compn of whose chloroplatinate C₁₁H₁₄O₂N₂Cl₂Pt, yellow, m. 204°, shows that the HO group has remained intact. Me ether of XXI, b₁₁ 120–5°, picrate, m. 155°, chloroplatinate, m. 172°; methiodide, m. 115° (effervescence) (the corresponding derivs of XXI m. 157°, 180°, 147°, resp). Methiodide of XXIII, m. 172° (foaming). XXIV, from 1% H₂SO₄NNaMe and Cl(CH₂)₃OH on the H₂O bath, b₁₁ 194°, b₁₂ 167°. Ephedrine with Me₂SO readily yields the quaternary salt of the Me ether; the corresponding iodide, C₁₁H₁₇ON₂I, m. 190°, smoothly loses MeI when heated *in vacuo*, yielding the Me ether, b₁₁ 110°; picrate, m. 106–8°. A methoxy β bromodihydrosafrole (XXXV), obtained from isosafrole dibromide (XXXVI) with MeOH, and which certainly has the MeO group on the C atom, gives with Me₂SO, and then KI an unusually hygroscopic quaternary iodide (XXXVII), yielding a chloroplatinate, m. 216° while the hydroxybromoisosafrole from XXXVI and H₂O yields, together with a little XXXVII, chiefly the non hygroscopic methiodide, m. 190°, of XXXIV; chloroplatinate, m. 195°. β Hydroxydiethyl sulfone (XXVI), from EtSO₃Na and ClCH₂CH₂OH in alc on the H₂O bath, b₁₁ 193–5°, m. 46°; Me ether, b₁₁ 142–4°. Dimethyl β hydroxyethylsulfonium iodide (XXVII), from MeSCH₂CH₂OH and MeI, m. 60–2°, absorbs moisture from the air (chloroplatinate, red brown, m. 80–1°). γ Hydroxypropyl homolog (XXVIII), m. 52–5° (chloroplatinate, orange, m. 115°). XXX with Me₂SO, gave a small quantity (1.5 g from 10 g XXX) of a compd, b₁₁ 95–6°, having the compn of the Me ether of XXX. The Me ether of XXX, m. 71°, b₁₁ 121–3°. Me ether of XXXI, b₁₁ 134–6°, m. 62°, chloride, b₁₁ 85°, reacts readily with C₆H₅I and AlCl₃ to give the ketone PhCH(CH₂OMe)COPh, yellowish, b₁₁ 139–41°, converted by coned HCl into the ketone PhCH(CH₂Cl)COPh, light yellow, b₁₁ 139–40° with some loss of HCl. XXXII gives the acid o HO₂CC₆H₄CH₂OMe, m. 93–4°, b₁₁ 121–5°, does not form a chloride with SOCl₂ but loses MeCl with regeneration of XXXII. XXXIII gives a compd b₁₁ about 155°, m. 95–6°, having the compn of (α methoxybenzyl) succinic anhydride, and giving the acid m. 140°, when pptd from Na₂CO₃ with acids.

C. A. R.

Derivatives of chloroacetylated phenols and phenol-alcohols. A. LAEGER. *Maz-yar Chem. Folyoisrat* 35, 151–6 (1929). —ICH₂CO derivs of di-HO derivs of C₆H₄ like ClCH₂CO derivs, are hydrolyzed by NH₃. With AgNO₃ halogenoacetylphenols react vigorously at 96°, affording dimro derivs and Ag halogenoacetate, AgCN behaves analogously. The following are described: chloroacetates of pyrocatechol, m. 50°, b₁₁ 185–6°, resorcinol, m. 76°, b₁₁ 191–2°, quinol, m. 128°, iodacetates, m. 48–0°, 50–60°, 112–3°, resp. Quinol chloroacetate forms with PhNMe₂ an equimol compd, m. 152–3°, decomp 156–7°. Phenylethyl aminoacetate, m. 115°, phenylpropyl aminoacetate, m. 105°.

B. C. A.

Propargyl ethers of phenol. CHARLES D. HURD AND FRANK L. COHEN. *J. Am. Chem. Soc.* 53, 1068–77 (1931). —PhC≡CBr (I) was obtained from PhC≡CCOCl and EtOH, giving the Et ester and reacting this with PhMgBr. The yield of the ester is 65% and of the ketone 85%, the ketone b₁₁ 200–2° and m. 53–5°. PhC≡CCPh₂OH results in 65–70% yield from BrPh and PhC≡CNa. MeC≡CH and EtMgBr give a Grignard reagent which reacts with Me₂CO to give 47% of 2 methyl-3-penten-2-ol (trimethyl propargyl alc) (II), b₁₁ 75–7°. n_D²⁰ 1.4193, a by-product is 2 methyl-1-penten-3-ene.

$\text{C}_{11}\text{H}_{17}\text{CMeC-CMe}$ (III), b. $75-7^\circ$, n_D^{20} 1.4002. I-Phenylethynylcyclohexan-1-ol, b₁ $166-9^\circ$, m. $58-60^\circ$. PhC CNa and $p\text{-Me}_2\text{NC}_6\text{H}_4\text{Bz}$ give 50% of α,γ -diphenyl- α - p -dimethylaminophenylpropargyl alc, m $144-5^\circ$. II gives a chloride, light yellow, b₁ $57-61^\circ$, n_D^{20} 1.4143 (33% yield) γ -Methylpropargyl chloride (1-chloro-2-butene), light yellow, b $81-4^\circ$ (14% yield) Triphenylpropargyl Ph ether (IV), from the chloride and PhONa in 70% yield, m $90-90.5^\circ$, trimethylpropargyl Ph ether (V), viscous yellow oil, n_D^{20} 1.3408, γ -methylpropargyl Ph ether (VI), b₁ $123-6^\circ$, n_D^{20} 1.3894 (57% yield) IV, heated slowly and uniformly to 310° , gives PhOH and 9 phenylethynylfluorene, yellow, m $98-100^\circ$, this was also prepd from 9-chlorofluorene and PhC CNa Both V and II decompose on heating into PhOH and III. When refluxed, VI changed primarily into a tar, the apparent mol wt. of which was 400 No simple allenes were found although they may have been the precursors of the tars

C. J. WEST

Bromine derivatives of certain mixed ethers and some of their reactions. L. CHAS. RAIFORD AND LOUIS H. HOWLAND *J. Am. Chem. Soc.* 53, 1051-7 (1931).—Pentabromophenyl alkyl ethers *iso*-Pr, m 86° (nearly quant yield), Bu, m $79-80^\circ$ (quant yield), *iso*-Bu, m $92-3^\circ$ (45% yield), *sec*-Bu, m $57-8^\circ$ (25% yield), *iso*-Am, m $64-5^\circ$ (quant yield), no reaction was obtained with *tert*-Bu or α -Am halides. Four methods were studied for splitting the ethers (1) glacial AcOH soln. of ether and PBr_3 , to which the calcd amt. of H_2O was gradually added was heated under a reflux, (2) the mixt in (1) was heated in a sealed tube at $135-50^\circ$ for 12-15 hrs., (3) Br was dropped into a C_6H_6 soln. of the ether contg I as a catalyst, (4) the ether, excess Br and about 1% Al as AlBr_3 , reacted at room temp. for 2 days. Of the tri Br derivs of the Ph alkyl ethers, only *iso*-Pr and *sec*-Bu were completely split by (1), these were also split by (3) (others not studied), Me, Et, Pr, *iso*-Bu and *iso*-Am were completely split by (2) but showed no reaction with (1). Of the penta-Br derivs, *iso*-Pr and *sec*-Bu were completely split by (1), Me, Et, Pr, Bu, *iso*-Bu and *iso*-Am were completely split by (2), Me, Et and *iso*-Pr showed no reaction with (3), the following % of Br_2COH were obtained in (4), Me 80; Et 40, Pr 40, *iso*-Pr 60, Bu 30, *iso*-Bu 20, *sec*-Bu 70, *iso*-Am 30. 2,4-Dichlorophenyl alkyl ether (I), b₁ $144-5^\circ$, 3,5-di-Br deriv (II), b₁ 145° , 2,6-di-Br deriv (III), b₁ $132-3^\circ$; penta Br deriv (IV), m $167-8^\circ$. I and Br in CHCl_3 at 0° give the β,γ -di-Br deriv, b₁ 188° (95% yield), II gives the 2,3,4,5-tetra-Br β,γ -di Br deriv, m $123-4^\circ$ (97%); III gives the β,γ -di-Br deriv, m $48-9^\circ$, as does IV, m $122-3^\circ$ (96 and 94%, resp.). Alc. KOH transforms these into derivs of Ph propargyl ethers 2,4-di Cl, liquid, decomps on distn *in vacuo*, 2,4-di-Br, m 65° , 2,6-di-Br, m $58-60^\circ$, 2,4,6-tri-Br, m $136-7^\circ$, the yields are 97-9%. The Ag salts (with 1 mol. AgNO_3) of these ethers were prepd, as well as the Cu and Hg salts of the 2,4-di-Br and 2,4,6-tri-Br derivs. The HBr evolved in the bromination of Ph alkyl ethers may split them unless AcONa is present. The splitting is most complete when the alkyl group is connected through a *sec*, or *tert*, C atom. Ph alkyl ethers contg 2 or more Br atoms in the nucleus could not be rearranged by heat

C. J. WEST

Effect of substituents on certain physical properties of benzene picrate. O. L. BAREL AND E. S. HAUBER. *J. Am. Chem. Soc.* 53, 1087-91 (1931).—The addn of HO groups to benzene picrate (I) lowers the m p and deepens the color of the picrate formed, the addn of Me groups to I raises the m p of the product and deepens the color of the picrate less than the HO group. The addn of Me groups to a side chain of I gradually increases the m p, and deepens the color of the picrate, this increase is more rapid than when a corresponding no. of Me groups are substituted in the ring. The sym tri-HO and tri-Me derivs. of I are the darkest and the color of the picrates become lighter as we approach the 1,2,3-derivs. The same deepening of color was obtained with the Me derivs. of C_6H_5 . Unsatur in the side chain yields a very unstable picrate. In the following list of picrates the color and m p are given, C_6H_5 , 83.9° , PhMe, light yellow, 88.2° ; PhEt, light yellow, 96.6° ; PhPr, yellow, 103.5° , *o*-, *m*-, and *p*- $\text{C}_6\text{H}_4\text{Me}$, lemon-yellow, 88.5 , 90.15° and 90.5° , resp., 1,2,3- $\text{C}_6\text{H}_3\text{Me}$, yellow, 89.5° , 1,2,4- $\text{C}_6\text{H}_3\text{Me}$, yellow, $96-7^\circ$, 1,3,5- $\text{C}_6\text{H}_3\text{Me}$, bright yellow, 96.6° , 1,2,3,4- $\text{C}_6\text{H}_2\text{Me}$, golden yellow, $92-5^\circ$, $\text{C}_6\text{H}_4\text{Me}$, golden yellow, 131° , C_6Me_6 , orange-yellow, 170° , all except the last 3 of these are unstable, decompose in the air. PhOH, yellow, m 83.1° , *o*-, *m*- and *p*- $\text{C}_6\text{H}_4\text{OH}$, orange-yellow (*o*- and *m*-) and light yellow, m 122° , $89-90^\circ$ and $115-7^\circ$, resp., 1,2,3- $\text{C}_6\text{H}_3\text{OH}$, lemon-yellow, $128-9^\circ$, 1,2,4- $\text{C}_6\text{H}_3\text{OH}$, orange-red, 96° , 1,3,5- $\text{C}_6\text{H}_3\text{OH}$, brown, $101-3^\circ$, *o*-, *m*- and *p*- $\text{MeC}_6\text{H}_4\text{OH}$, orange-yellow, yellow and bright red, m 88° , 61.6° and $64-5^\circ$, resp., *o*-cresol, orange-yellow, 92° , 1,2,4-, 1,3,2- and 1,4,2- $\text{Me}_2\text{C}_6\text{H}_3\text{OH}$, chrome-yellow, orange-yellow and orange, m 83.8° , $50-3^\circ$ and $81-2^\circ$, resp., C_6H_5 , yellow, 149° ; α - and β -Me derivs, lemon-yellow, $140-1^\circ$ and $115-6^\circ$, resp.; α - and β -Et derivs, lemon-yellow and yellow, m 97.4 and $69-70^\circ$, resp.; α - and

β Pr derivs., yellow, m 140-1° and 89-90°, resp; α - and β Bu derivs., orange-yellow, m 104-5° and 71-3°, resp. 1,4 and 2,6-di Me derivs., orange-yellow, m 139-40° and 141-2°, resp. 1,2,6 tri Me deriv., orange yellow, m 121-2°, α - and β C₁₁H₁₇OH, orange-yellow, m 185.5° and 155.5°, resp., anthracene, ruby red, m. 138°; hexa-Me deriv., dark brown, m 203°, phenanthrene, orange-yellow, m 132.8°, acenaphthene, orange red, m 160.5°, stilbene, red, m 90-1°, fluorene, red brown, m. 77°, pyrene, red m 220°

Phenyl ethers MANFRED OESTERLIN *Monatsh* 57, 31-44(1931)—PhONa and p -BrC₆H₄OMe with Cu at 201-10° for 4 hrs give p -methoxydiphenyl ether, b₁₁ 163-5°, AlCl₃ in C₆H₆ splits off the Me group, giving the p -HO deriv (I), m 84°, bromination (KBr and KBrO₃ with concd H₂SO₄ in CS₂) gives 82% of the p -Br deriv (II), m 85° I in KOH and p -BrC₆H₄OMe with Cu at 210° for 4 hrs. give p -methoxy- p' -phenoxydiphenyl ether, m 82° AlCl₃ gives the p -HO deriv, m 87°. p -MeOC₆H₄OH and p -BrC₆H₄OMe with Cu and 80% KOH give 65% of p - p' -dimethoxydiphenyl ether, m 102°, the di-HO deriv, m 160°. p -MeOC₆H₄OH and p -C₆H₄Br with KOH and Cu at 200-10° for 5 hrs. give di(p -methoxyphenyl) hydroquinone ether (p -MeOC₆H₄O)₂C₆H₄, this also results from II and p -MeOC₆H₄OH with Cu and KOH at 210° for 4 hrs. The di p -HO deriv, m 188°. p -MeOC₆H₄OH and (p -BrC₆H₄)₂O give p -di(methoxyphenyl)dioxydiphenyl ether, (MeOC₆H₄OC₆H₄)₂O, m 164°. p -O₂C₆H₄OPh, m. 61°, NH₃ deriv, m 84°, Ac deriv, m 131° nitration gives p -acetylaminom-nitrodiphenyl ether, yellow, m 104°, reduction gives the m NH₂ deriv, m 124°, di NH₂ deriv, m 66°. p -Amino- m -nitrodiphenyl ether, red-orange, m 47-8°, then solidifies and m 81°, the reduction product is the di NH₂ deriv, m 66°, HCl salt, m 216°, 3,3-dimethylbenzil gives 7,3-di-(p -nitrophenyl)-6 phenoxyquinoraline, light yellow, m. 195-6°. p -ClC₆H₄NO₂ and p -HOC₆H₄OPh with Cu and KOH at 150° for 5 hrs give p -phenoxy- p' -nitrodiphenyl ether, yellow, m 94°, p -NH₂ deriv, m 84°, Ac deriv, m 145°, nitration gives p -phenoxy- p' -acetamidom-nitrodiphenyl ether, m 121° sapon and reduction give p -phenoxy- m' , p' -diaminodiphenyl ether, m. 95°. p -MeOC₆H₄OH and p -ClC₆H₄NO₂ give p -methoxy- p' -nitrophenyl ether, yellow, m 111-2°, p' -NH₂ deriv, m 81-2°, HCl salt, m 212°, sulfate, m 220°, Ac deriv, m 131°, nitration gives the m' -nitro deriv, m 106°, p -methoxy- m' -nitro- p' -aminodiphenyl ether, m 70-7°, m' , p' -di-NH₂ deriv, m 105°, p -Di-(nitrophenyl)dioxydiphenyl ether, m 130°, di NH₂ deriv, m 109°, di Ac deriv, m 205-6°.

C. J. WEST

Diphenyl ether series. II. Preparation and structure of some sulfonic acids and related derivatives. C. M. SUTCH, *J. Am. Chem. Soc.* 53, 1112-6(1931), cf. C. A. 23, 4460—Bromination of Ba phenoxybenzene-4,4'-disulfonate and treatment of the ppt with Na₂CO₃ soln gives 57% of the Na 4-bromophenoxybenzene 4'-sulfonate, p -toluidine salt m 245-7°. Further bromination of the Na salt gives 84% of (4-BrC₆H₄)₂O Ph₂O and Ac₂O, shaken with 95% H₂SO₄ and heated 1 hr on the H₂O bath give 93% of the mono SO₃H acid (cf. Quilico, C. A. 22, 1703), p -toluidine salt, m 221-2°, bromination of the Na salt gives (4-BrC₆H₄)₂O, fusion of the Na salt with NaOH gives little or no 4 HOC₆H₄OPh. With PCl₅, heating 10 hrs. at 170-80° gives 54% of 4 phenoxybenzenesulfonyl chloride, b₁₁ 205-7°, m 45-6°, amide, m 128-9°, anilide, m 86-8°. 4 BrC₆H₄OPh and 95% H₂SO₄ give a mono-sulfonate, ClSO₂H gives 4 bromophenoxybenzene 4'-sulfonyl chloride, m 81-2°, also obtained in 91% yield by bromination of 4 PhOC₆H₄SO₂Cl, amide, m 131-2°, anilide, m 108-9°. Phenoxybenzene-4,4'-disulfonyl chloride m 128-9°, diamide, m 158-60°. Reduction of the sulfonyl chloride with Na₂SO₃ and Na₂CO₃ gives 4 phenoxybenzenesulfonic acid, m 100-1°, which drops to 91-3° in 48 hrs. Reduction of the chloride with Zn and H₂SO₄ gives 60% of 4 thiodiphenyl ether, b₁₁ 178-80°, m 21-2°, oxidation of the Na salt with air or with I in EtOH gives the disulfide m 47-8°.

C. J. WEST

3,5-Dimethoxyaniline. REINHARD SEKA AND WALTER FUCHS *Monatsh* 57, 63-70(1931)—3,5(MeO)₂C₆H₃CO₂Me (I), m 43°, and liquid NH₃ in a bomb tube at room temp for 4 weeks give 92% of the amide, m 149-9°, which may also be prepd from the acid chloride and NH₃, the Hoffmann degradation gives about 52% of 3,5 dimethoxyaniline, b₁₁ 177°, m 46°, HCl salt, m 210°, chloroplatinate, golden yellow, decomp 223°, picrate, yellow, decomp 167-70°, Bz deriv, m 139°, anisoyl deriv, m 119-20°. I and N₂H₄, H₂O, boiled 8 hrs, give 50% of the hydrazide, m 168.5° (HCl salt m 210°), the azide, m 50-1°, decomp about 70° and with H₂SO₄ gives (MeO)₂C₆H₃CO₂H, urethan m 43.5°.

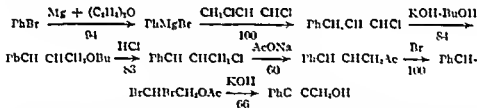
C. J. WEST

The behavior of p -acetophenetide toward chlorine and bromine. E. BURAS AND J. KOVÁŘOVICOVÁ, *Časopis Českosloven Lékárnictva* 10, 197-202, 233-9(1930)—Chlorination of p -acetophenetide (I) in glacial AcOH at ordinary temp and pressure gave 2,3,5,6 tetrachloro- p -acetophenetide (II), m 226°, while bromination of I resulted in 3,5-di

bromo-*p*-acetophenetide (III), *m* 175°. The entrance of the halogens caused an increase in stability, and a decrease in basicity and in ease of replacement of NH₂. Sapon of II gave 2,3,5,6-tetrachloro-*p*-phenetide, *m* 96°, from which the following ethoxybenzenes were prepd: 2,3,5,6-tetrachloro, *m* 56°; 2,3,4,5,6-penta-Cl, *m* 75°; 2,3,5,6-tetrachloro-4-bromo, *m* 81°; 2,3,5,6-tetrachloro-4-iodo, *m* 79° and 2,3,5,6-tetrachloro-*p*-quinone. Sapon of III yielded 3,5-dibromo-*p*-phenetide (IV), *m* 54°, the constitution of which was detd by conversion to 3,5-dibromophenetole (V), *b* 267°. Attempts to replace the NH₂ group in IV by OH or CN were unsuccessful, but the following 3,5-dibromoethoxybenzenes were prepd: Me₂N, *m* 84°, BrNH, *m* 159°, EtNH, *m* 91°, Ag₂N, *m* 108°; 4 Br *m* 72°; 4 I *m* 91°; 4 Cl, *m* 42-4°, and the picrate of IV, *m* 145°. Nitration of III yielded 3-bromo-5-nitro-*p*-phenetide, *m* 108°. Nitration of V gave 3,5-dibromo-1-ethoxy-2-nitrobenzene, *m* 125°, which on reduction yielded IV. 3,4,5-Tribromo-1-ethoxy-2-nitrobenzene, *m* 32-3°, was prepd. WILLIAM J. HUGA

Dehydration with potassium hydroxide of the ethylol group adjoining the benzene nucleus. III. Dehydration of the α,α -alkylarylethanol. SEBASTIEN SABETAY. *Bull. soc. chim.* [4], 47, 614 (1930) of C. A. 24, 1099.—In the distn of alcs. of the phenylethyl alc type over anhyd. KOH the secondary and tertiary alcs. are much more stable than the primary ones that easily form the corresponding unsatd. hydrocarbon. S. shows that of *hydratropic alc.*, PhCHMeCH₂OH (I), and *methylethylcarbinol*, PhCH₂CHMeOH (II), only I is dehydrated, to α -methylstyrene, PhCH=CH₂ (III), by anhyd. KOH while *dehydration with poronic* (Ramat Lucas and Amagat, C. A. 22, 1582, and Fournau, C. A. 16, 2195) forms a mixt. of III and PhCH=CHMe. By heating 10 g. of I over 5–7 g. anhyd. KOH 2 layers form, the upper one of which is distd. off, extd. with Et₂O, dried and rectified, finally over Na. III b.p. 159–60°, n_D^{20} 1.5300, d_4^{20} 0.9065, mol. refraction 40.56. It tends less to polymerization than styrene and forms a liquid *di bromide*. The identity of III was established by comparing it with III prepd from PhMe₂COH according to Staudinger and Breusch (C. A. 23, 3213) as well as with "poronic" (cf. references above). α -Ethylstyrene, PhEtCH=CH₂ (IV), was similarly prepd by the action of anhyd. KOH on PhEtCHCH₂OH. IV b.p. 177–80°, n_D^{20} 1.5282. Its *di bromide* seems to be liquid. Dehydration of II under the same conditions yields only traces of PhCH=CHMe. G. TOENIES.

New method of synthesis of phenylpropargyl alcohol and its homologs substituted in the ring. L. BARR *Compt. rend.* 191, 493-5 (1930).—Instead of condensing $\text{PhC}\equiv\text{CH}$ with HCHO , which will give very poor yields, B prepd $\text{PhC}\equiv\text{CCl}_2\text{H}$ in the following way (the figures under the arrows indicate the % yield obtained)



The com cinnamyl acetate, which is employed in perfumery, can be used as a starting material

Reduction of Schiff bases. L. ZECHMEISTER AND J. TRUKA, *Ber.* 63B, 2283-4 (1930).—Mg in MeOH can be used advantageously for the reduction of the $-CH=N-$ group in Schiff bases to $-CH_2NH-$. The reaction takes place in a neutral soln. and even when the product is decomposed with H_2O there is no marked alkyl. The base is dissolved in almost anhyd. MeOH, an at least 5-fold excess of Mg ribbon is added (the reaction being regulated by cooling or heating, as needed), then the MeOH is distilled off, the residue decomposed with ice H_2O and the $Mg(OH)_2$ dissolved out with 30-50% AcOH. The following amines were obtained in 70-90% yield in this way: PhCH₂NHPh, m. 37-8°, β -MeOC₆H₄CH₂NHPh, m. 64.5°, α -HOC₆H₄CH₂NHPh, m. 108°, PhCH₂CH₂NHPh, m. 193-5°, 3,4-Cl₂O₂C₆H₂CH₂NHPh, m. 78°, β -Me₂NC₆H₄CH₂NHPh, m. 62°, β -MeC₆H₄NHCH₂Ph, m. 31s-20°, α -C₆H₅CH₂NHCH₂Ph, m. 67°, α -C₆H₅CH₂NHCH₂CH₂OMe(p), m. 80°, β -C₆H₅CH₂NHCH₂Ph, m. 68°, β -C₆H₅CH₂NHCH₂C₆H₄OMe(p), m. 104.5°.

Acyl derivatives of *o*-aminophenol. VI. C. B. POLLARD AND R. E. NELSON, *J. Am. Chem. Soc.*, 53, 996-1001 (1931). cf. C. J., 23, 4940—A study of the diacyl derivs. of *o*-H₂NC₆H₄OH, when 1 of the acyl groups was always the PhCH₂CO radical.

Is reported. When both acylating agents were of the type CICOR, isomeric diacetyls were obtained depending on the order of the introduction of the acyl groups. In 5 cases out of 8 complete rearrangement occurred during sapon. 1 case showed partial rearrangement and in 2 cases rearrangement did not occur. Apparently relative acidity and wt. are not the controlling factors in this type of rearrangement. When 1 of the acylating agents was an alkyl chlorocarbonate, the same diacyl deriv was obtained regardless of the order of introduction of the groups. *o*-Phenylacetylaminophenol (I), m 149-50° (60% yield), *o*-butyrylamino deriv, m 80-1° (58% yield); *o*-(*m*-bromobenzoylamino) deriv, m 180° (60% yield), *m* Cl isomer, m 156-8° (52% yield). 1 and AcCl with a drop of H₂SO₄ give 60% of the acetate, m 101-2°; *o*-AcNH₂CH₂OH and PhCH₂COCl give 45% of *o*-acetylaminophenyl phenylacetate, m 99-100°, sapon. of either gives PhCH₂CONHCH₂OH, showing that rearrangement occurred in the second case. *o*-Phenylacetylaminophenyl propionate (II), from 1 and EtOCOCl, m. 99-9° (40% yield), butyrate (III), m 91-2° (40% yield), valerate (IV), m 80-2° (50% yield), isovalerate (V), m 87-8° (50% yield), chloroacetate (VI), m. 106-7° (50% yield); benzoate (VII), m 110-1° (60% yield), *m*-chlorobenzoate (VIII), m. 146-8° (65% yield). *o*-Propionylaminophenyl phenylacetate, m 71-2° (35% yield), sapon. of this or of II gives PhCH₂CONHCH₂OH (IX), the *o*-butyrylamino deriv, m 40-8° (40% yield), gives PhCONHCH₂OH on sapon, while the isomeric III gives IX; the *o*-valeryl amino deriv, m 71-2° (40% yield), gives BuCONHCH₂OH (X) on sapon, while the isomeric IV gives 65% IX and 35% X; *o*-isovaleryl amino deriv, m 50-7° (55% yield), gives iso-BuCONHCH₂OH, as does V, on sapon; *o*-chloroacetyl amino deriv, m 113-4° (55% yield), or VI yields IX on saponification, *o*-benzoylamino deriv, m 108-9° (75% yield), yields BzNHCH₂OH on hydrolysis, while the isomeric VII gives IX; *o*-(*m*-chlorobenzoylamino) deriv, m. 150-2° (55% yield), or VIII give on hydrolysis ClCH₂CONHCH₂OH; *o*-(*m*-bromobenzoylamino) deriv, m 157-9° (55% yield), gives the corresponding Br deriv. on hydrolysis. The phenylacetate of *l*-*o*-hydroxycarbamide, m 105-6° (60% yield), yields on hydrolysis 50% each of IX and MeOOCNHCH₂OH, the *Et* deriv, m 62-3° (40% yield), gives EtOOCNHCH₂OH, the *iso*-Bu deriv, m. 72-3° (53% yield), gives the *iso*-Bu deriv.

C. J. WEST

Reduction potentials of some higher benzologs of the quinones. LOUIS F. FIESER AND EMMA M. DIERZ. *J. Am. Chem. Soc.* 53, 1128-33 (1931).—A potentiometric study of the benzologs of anthraquinone and phenanthrenequinone leads to the following conclusions. The potentials of the benz- and dibenzanthraquinones lend support to the *o* quinonoid theory of the structure of anthracene. Hydroxychrysenequinone constitutes an exception to the rule that when a quinone can exist in 2 forms, the 1 of lower potential will predominate. The following values for the reduction potential at 25° (*E*₀ (av.) in v) are reported: 1,2-benzanthraquinone, 0.223, 1,2,7,8-dibenzanthraquinone, 0.264, 1,2,5,6-dibenzanthraquinone, 0.268, (1,2)Bz-3-metho-1,2,5,6-dibenzanthraquinone, 0.257, 5,6-chrysenequinone, 0.465, 6,12-chrysenequinone, 0.393, 12-hydroxy 5,6-chrysenequinone, 0.391, 12-EtO deriv, 0.418, 1,2-benz-3,5-anthraquinone, 0.430, picenequinone, 0.474.

C. J. WEST

Configurational relationships of phenylated carbimins. III. P. A. LEVENS AND A. WALT. *J. Biol. Chem.* 90, 81-8 (1931), cf. *C. A.* 24, 4008.—A direct chem. method is found for the correlation of the configuration of Me(PhCH₂)C(OH) with sec aliphatic alcs. The method is based on the condensation of optically active ethylene oxides with aliphatic or aromatic radicals by the Grignard reaction. Condensation of *d*-MeCH(CH₃)O (I) with C₆H₅MgBr gave *l*-MeCH(OH)C₆H₅, [α]_D²⁵ -1.75° (without

solvent), and its *o*-naphthylurethan gave [α]_D²⁵ -4.23° (in abs. alc.). PhMgBr and I gave *l*-Me(PhCH₂)C(OH), [α]_D²⁵ -19.8° (without solvent); *o*-naphthylurethan, [α]_D²⁵ -31.6°. The latter carbinol was oxidized to the corresponding ketone which was then identified by the semicarbazone, m 188°. L. and W. conclude that the direction of rotation is detd by the position of the heavier group, provided the Ph group is not less than 1 C atom from the asym. C atom.

H. W. LEAHY

Beckmann rearrangement. IV. Unsaturated ketoximes. KARL V. AUWERS AND MARIA SEYFRIED. *Ann.* 484, 178-211 (1930); cf. *C. A.* 19, 1405; 23, 4701.—PhCHClIBz (I) with NH₂OH in alk. soln. gives 3,5-diphenylisoxazoline, m. 75°, and a mixt. of alkali sol. hydroxyaminoximes. The oxime (II) of I, m 115-6°, was prepd by the use of NH₂OH.HCl; this was unchanged by boiling with 2 N NaOH for 10 hrs. The Beckmann rearrangement gives PhCHClCHCONHPh (cf. Henrich, *C. A.* 1, 1411). II yields a dibromide, m 156°, which splits off IIBr when heated above its m. p. giving 3,5-diphenylisoxazole (III). Beckmann rearrangement gives a dibromohydrocinamic

anilide, m 179°. PhCHBrCHBrBz could not be transformed into the oxime $\text{PhCH}=\text{CBrBz}$ yields an oxime, m 151° (C. A. 23, 4701), which gives III with EtOH-KOH . The Beckmann rearrangement product, probably $\text{PhCH}=\text{CBrCONHPh}$, was obtained only as a brown smeary product $p\text{-MeOC}_6\text{H}_4\text{COCH CHPh}$ (IV) and PhNH_2 in AcOH give 1,5-diphenyl-3-p-anisylpyrazoline, m 140.5–1.5°, the concd H_2SO_4 soln becomes blue-violet on addn of FeCl_3 , NH_4OH in alk soln gives 3-p-anisyl-5-phenylisoxazoline, m 104.5–5.5°, oxidized by CrO_3 in AcOH to the isoxazole (V), m 121°, this is also obtained from $p\text{-MeOC}_6\text{H}_4\text{COC}(\text{C}_6\text{H}_5)_2$ and $\text{NH}_4\text{OH HCl}$, while free NH_4OH gives 3-phenyl-5-anisylisoxazole, m 127–8°. IV and $\text{NH}_4\text{OH HCl}$ in EtOH give benzal $p\text{-methoxyacetophenone oxime}$, m 140–2°, Ac deriv , m 134–5°. Beckmann rearrangement gives cinnamic $p\text{-aniside}$, m 152–3°. $p\text{-MeOC}_6\text{H}_4\text{COCBr CHPh}$, b_p 250–3°, with NH_4OH in NaOH gives V, $\text{NH}_4\text{OH HCl}$ gives the oxime, m 153°. The relative rates of decompn of the ketone and oxime by EtOH-KOH are 10 min, 17, 86%, 20 min, 50, 95%, 30 min, 66, 95%, 15 hrs, 100, 95%. Beckmann rearrangement gives a brown smear, from which $p\text{-MeOC}_6\text{H}_4\text{COH}$ was isolated after hydrolysis β Bromobenzal $p\text{-methoxyacetophenone}$, from $\text{PhC}:\text{CCOC}_6\text{H}_4\text{OMe}$, yellow, m 65–6°, NH_4OH and alkali give V; $\text{NH}_4\text{OH HCl}$ gives 50% of the oxime (VI), m 115–6°. The relative rates of decompn of the ketone and oxime by alkali are 10 min, 10, 11%, 20 min, 93, 51%, 30 min, —, 68%, 15 hrs, 93, 85%. The relative rates of decompn of the α and β -bromo oximes with 2% NaOH are 2 min, 10, 11%, 7 min, 45, 14%, 12 min, 70, 10.5%, 17 min, 93, 10.5%. The Beckmann rearrangement of VI gives β -bromocinnamic $p\text{-aniside}$, m 144°. Benzal $p\text{-ethoxyacetophenone dibromide}$, m 151–2°. V. Substituted cinnamaloximes and cinnamonitriles. *Ibid* 212–35 — $\text{cis-PhCH}=\text{CBrCHO}$ (I) is oxidized to $\text{trans PhCH}=\text{CBrCO}_2\text{H}$ by alk. Ag_2O or by O_3 in CHCl_3 . The oxime (II), m 144°, results in acid, alk or neutral medium, warming with $\text{EtOH H}_2\text{O KOH}$ gives $\text{PhC}:\text{CCN NOH}$, m 108°, Ac deriv of II, m 81–2°, heating II with AcO gives $\text{cis-PhCH}=\text{CBrCN}$ (III), b_p 138–40°, d_4^{25} 1.4624, n_D^{25} 1.61084, 1.62083, 1.63069 for α , He and β , this also results from the acid amide and POCl_3 at 70–2°. *Dibromocinnamonitrile*, m 92–3°, also gives III with EtOH-KOH . II and PCl_5 in dry Et_2O , allowed to stand 10–24 hrs, give the *trans* isomer (IV), b_p 148–50°, m 33–5°, d_4^{25} 1.4263, n_D^{25} 1.60386, 1.61348, 1.63004 for α , He and β , IV was also obtained from the *trans* acid amide, m 120–1°, and POCl_3 at 70–2°. Attempts to change III to IV or IV to III failed. The relative rates of reaction of III and IV with EtOH-KOH are 2 min, 42.3, 78% decompose, 10 min, 99.6, 42.6% III or IV and 10% EtOH KOH give β -ethoxycinnamonitrile, b_p 146–8°, d_4^{25} 1.0601, n_D^{25} 1.55988, 1.56718, 1.58641, 1.60482 for α , He , β and γ . Addn of EtOH to $\text{PhC}:\text{CCN}$ gives an isomeric nitrile, b_p 163–8°, d_4^{25} 1.0543, 1.54569, 1.55169, 1.56736 for α , He and β . Both forms give BzCH_2CN with dil H_2SO_4 and show little difference in their reaction with Br in CS_2 . $\text{PhC}:\text{CCHO}$ and HBr in AcOH give β -bromocinnamaldehyde, b_p 144–6°, d_4^{25} 1.492, d_4^{25} 1.4937, n_D^{25} 1.62865, 1.63884, 1.66649 for α , He and β , it slowly decomposes in the air, giving off HBr . Heating with 1.1 mols AcOK in abs EtOH for 7 hrs. splits off 24.4% of the Br , *oxime*, m 81°, sepd by means of the Na salt into 2 isomers, m 63–6° and 103° (*Ac deriv*, m 68–9°). The relative rates of decompn of the aldehyde and oxime by 5% EtOH-KOH are 5 min, 30.5, 2.2% 10 min, 77.2, 4.4% C J WEST

Reduction of the mixture of aldehyde or ketone with nitrite. II. A. OGATA AND S. IITRANO. *J. Pharm Soc Japan* 50, 1141–52 (1930), German abstr. 147–8, cf. C. A. 24, 4758 — Reduction of the mixt. of equiv quantities of BzH and NaNO_2 in alc. AcOH or HCO_2H with excess Zn dust gave PhCH_2NH_2 , (PhCH_2) $_2\text{NH}$, α , β -diphenyl- α -hydroxy- β -benzylaminorothane (I) and NH_3 as end product. I, $\text{C}_{20}\text{H}_{19}\text{ON}$, m 151°, sol in hot alc., but insol in H_2O HCl salt, m 226–8°, Bz deriv , m 135–6°. The formation of *benzylamine* is due to the formation of an intermediate compd, *benzaloxime*, followed by reduction, while *dibenzylamine* is due to the reaction between PhCH_2NH_2 and unreacted BzH , followed by reduction F. I. NAKAMURA

Isomerization of the hydroxy aldehydes. IV. Transformation of benzylbromooacetaldehyde and benzylglycolaldehyde. S. DANILOV AND E. VENUS-DANILOVA. *Ber.* 63B, 2765–75; *J. Russ. Phys.-Chem. Soc.* 62, 1697–1711 (1930), cf. C. A. 24, 1093 — Benzylbromooacetaldehyde (I), from $\text{PhCH}_2\text{CH}_2\text{CHO}$ and Br in CS_2 , is an unstable oil readily resinifying and polymerizing on standing, but it yields a cryst *monohydrate*, m. 87.5–8°, which permitted of studying the conversion of I into *benzylglycolaldehyde* (II) and the acid (III). Heated with H_2O and freshly pptd BaCO_3 , I gives II in quite good yields, together with 4% $\text{PhCH}_2\text{CH}_2\text{CO}_2\text{H}$ (IV). With Ag_2O are obtained 20% II, 9.4% III, 5.6% of a condensation product (V) of I and 31% IV. PbO_2 yields chiefly halogen-contg condensation products of I and 22.6% IV. The structure of II is es-

tablished by its prepn, its derivs and its oxidation products. KMnO_4 in aq. $\text{C}_6\text{H}_5\text{N}$ gives PhCH_2CHO , BrOH and HCl . Its isomerization into the HIO ketone is accompanied by the formation of a diketone $\text{PhCH}_2\text{CH}(\text{OH})\text{CHO}$ (II) \rightarrow $\text{PhCH}(\text{OH})\text{COMe}$ (VI) + PhCOCOMe (VII). II, b_p 120-1°, m 51.5-2°, reacts with NH_4AgNO_3 , fuchsins, SO_2 and Fehling soln., mol wt in freezing C_6H_6 141.7, oxime, m 123°, semicarbazone, m 136.5°, phenyllosazone, yellow, m 137°, benzoyl, m 70°, mol wt in freezing C_6H_6 252.6. III, m 97.5°, mol wt in freezing C_6H_6 150. VII, b_p 113-5°, dioxime, m 239-0°. VI, b_p 130-2°, mol wt in freezing C_6H_6 155-61, semicarbazone, m 194°. With PhMgBr VI yields $\text{PhMeC}(\text{OH})\text{CH}(\text{OH})\text{Ph}$, m 103°, gives with concd H_2SO_4 a deep red color destroyed by H_2O and is oxidized by CrO_3 AcOH to PhCOMe and HCOOH . C. A. R.

New p -bromophenacyl esters. S. G. POWELL. *J. Am. Chem. Soc.* 53, 1172 (1931).—The following p -bromophenacyl esters were prepd: trimethylacetate, m 76.5°, isocaproate, m 77.3°, enanthate m 69.2°, isooheptylate, m 75.5°, pelargonate, m 63.5°.

C. J. West

Oximes of α,β -unsaturated ketones and the Beckmann rearrangement. A. H. BLATT, *J. Am. Chem. Soc.* 53, 1123-41 (1931).—Certain α,β -unsatd ketones, when treated with NH_4OH HCl , yield oximes of the general configuration $\text{RCH}=\text{CHCR}'-\text{NOH}$. These oximes undergo a *trans* shift in the Beckmann rearrangement and when dissolved in concd H_2SO_4 rearrange in a different manner to yield the isomeric isoxazolines $\text{RCH}=\text{CH}_2\text{CR}'\text{NO}$. The so-called stable oximes of these α,β -unsatd ketones

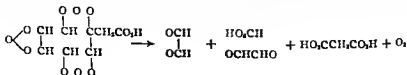
are in reality isoxazolines and are formed through a process which does not involve either the true oxime as an intermediate or the simple 1,4 addn of NH_4OH to the ketone. Benzal- o -chloroacetophenone oxime m 110-24° when heated at the rate of 2° per min, even though the oxime was pure, PCl_5 in excess in abs Et_2O gives *cinnamic* o -chloroanilide, m 136-7°, also obtained from the acid and o - $\text{ClC}_6\text{H}_4\text{NH}_2$. Boiling the oxime with MeOH-KOH for 5.5 hrs removes only 23% of the Cl . Benzal- p -bromoacetophenone oxime m 138-50°, there was no evidence of isoxazoline formation with LiOH-KOH or NH_4OH , PCl_5 gives the p -bromoanilide of cinnamic acid, m 191°, bromination of the oxime gives a β - Br deriv, m 155° (decompn); with EtOH-KOH it eliminates HBr , giving 3- p -bromophenyl-5-phenylisoxazole, m 178-9°; this also results in 50% yield by heating the oxime at 150° or by heating in glacial AcOH , it was also obtained from $\text{PhCH=CHBrCOC}_6\text{H}_4\text{Br}$ and NH_4OH HCl , the isoxazoline, m 138-9°, results from the oxime and concd H_2SO_4 or from the ketone, NH_4OH HCl and KOH .

C. J. West

Methods of manufacturing vanillin. J. SCHWYZER, *Chem-Ztg* 54, 817-8, 839-40 (1930).—Explicit details of the 4 practical methods for the manuf of vanillin are given. The methods are (a) chlorination of PhOH , treatment with $\text{Ba}(\text{OH})_2$ to form o - $\text{C}_6\text{H}_4(\text{OH})_2$, treatment of this with m - $\text{O}_2\text{NC}_6\text{H}_4\text{SO}_3\text{Na}$, CH_2O and Fe filings to give 3-4 $(\text{HO})_2\text{C}_6\text{H}_3\text{CHO}$ followed by methylation, (b) treatment of piperonal with PCl_5 and hydrolysis of the product to 3-4 $(\text{HO})_2\text{C}_6\text{H}_3\text{CHO}$, followed by methylation, (c) sulfonation of PhNO_2 , reduction by electrolysis or by Zn and HCl , and treatment with guaiacol FeSO_4 and CH_2O , or treatment of guaiacol with MeOH , $\text{Me}_2\text{NC}_6\text{H}_4\text{NO}$ and CH_2O , or guaiacol is treated with alloxan prepd by the action of Cl_2 on uric acid, sapon with NaOH , treatment with SO_2 , then with p - $\text{Me}_2\text{NC}_6\text{H}_4\text{NH}_2$, and finally with NaOH , (d) formation from clove oil.

C. H. Pratt

The action of ozone on benzoic acid and phenyl aliphatic acids. H. RUPP and H. HIRSCHMANN. *Helv. Chim. Acta* 14, 49-56 (1931).— BrOH and phenyl aliphatic acids add on a mol of O_3 at each of the double bonds of the benzene nucleus to form triozonides. These compds are white camphor like substances at low temps and oils at room temps. They are not very stable and tend to decomp. with explosive violence at room temp. They decomp with water to form, probably, 1 mol O_2 , 2 mols glyoxal, 1 mol HCO_2H and 1 mol of a dibasic acid. The yields of these hydrolysis products are poor. The yields of the ozonides are also poor as apparently much of the phenyl aliphatic acids are completely oxidized. The triozonide of $\text{PhCH}_2\text{CO}_2\text{H}$ (I) probably decomp with water as follows:



Ten g of I was dissolved in 100 cc CCl_4 and at $0-15^\circ$ treated with 10–12% O_2 for 24 hrs. The white ozonide was filtered and washed with ice-cold abs. ether and decompd. with water on the water bath. The clear soln gave 1.1 g of $(\text{Cl}_2\text{H}-\text{NNH}-\text{CO}-\text{NH})_2$ (IV), decompd. 300° . HCO_2H was identified by its odor and the sepn. of about 0.5 g which reduced 1.6 g Cl_2 and 1.6 g $\text{Cl}_2(\text{CO}_2\text{H})_2$ was identified by formation of $\text{PhCH}(\text{CHCl})\text{C}(\text{CO}_2\text{H})_2$, m $202-3^\circ$ instead of 208° as given by Rupe (*Ber* 37, 3123) and by the solid acid isolated in about 0.1 g yield, giving off CO_2 when melted. I (1.7 g) was recovered unchanged. In 1 expt 10 g of $\text{PhCH}_2\text{CH}_2\text{CO}_2\text{H}$ (II) in 100 cc. CCl_4 was treated at $0-15^\circ$ with 10–12% O_2 for 24 hrs. The white ozonide after decompn. with water yielded 4.2 g of IV, 0.8 g $(\text{Cl}_2\text{H}-\text{CO}_2\text{H})_2$, and only 0.1 g of II was recovered unchanged. Ten g of BrO_2H (III) in 80 cc CHCl_3 and 20 cc CCl_4 was treated at $0-15^\circ$ with 10–12% O_2 for 20 hrs. III gave the best yields of ozonide but this ozonide was the least stable. One sample of 3.5 g exploded in the desiccator. The ozonide from the above sample on decompn. with water yielded 0.62 g IV and 0.5 g $(\text{CO}_2\text{H})_2$, and 5.5 g III was recovered unchanged. III is more resistant to the action of O_2 than either I or II.

ODEN E. SHEPPARD

Condensation products of phenylacetic hydrazide, REINHARD SEKA AND ST. PETER HEILPERIN *Monatsh* 57, 45–51 (1931)— AcCO_2H and $\text{PhCH}_2\text{CONHNH}_2$ (I) in H_2O give 76.6% of the hydrazone, $\text{PhCH}_2\text{CONHNH}(\text{CMeCO}_2\text{H})_2$, m. 168° , the mother liquor contains diphenylacetic hydrazide (II), $(\text{PhCH}_2\text{CONHNH})_2$, m. $236-7^\circ$. With $\text{AcCH}_2\text{CO}_2\text{H}$ and I only II is formed. Levulinic acid gives the phenylacetic hydrazone, m. 119° ; the mother liquor contains II. Galactose likewise gives a hydrazone, m. $192-3^\circ$ (53% yield). Ac and I give a monophenylacetic hydrazone, m. 138° , sol. in H_2O , and a diphenylacetic hydrazone, $\text{C}_{15}\text{H}_{15}\text{O}_4\text{N}_4$, decompd. 254° . Benzil gives a diphenylacetic hydrazone, $\text{C}_{20}\text{H}_{15}\text{O}_4\text{N}_4$, m. $108-9^\circ$.

C. J. WEST

Several new 4'-sulfo-o-henzoylbenzoic acid derivatives and the corresponding anthraquinone compounds, IVAN GUBELMANN, H. J. WEILAND AND O. STALLMANN, *J. Am. Chem. Soc.* 53, 1033–6 (1931)— $2 \text{HO}_2\text{C}_6\text{H}_4\text{COC}_6\text{H}_4\text{SO}_3\text{Na}$, m. 148° , nitrated with a mixed acid (100 g mixed acid contg. 70% HNO_3 , 22% H_2SO_4 , and 8% H_2O with 480 parts of 25% oleum) in 100% H_2SO_4 , the acid being added over 6–8 hrs at $15-20^\circ$ and the reaction then heated at 35° for 3 hrs, gives the 5'- NO_2 deriv.; 100 g of the free acid, heated with 25% oleum at 150° , give 80 g of 2-nitroanthraquinone-7-sulfonic acid, with NaClO_2 in dil. HCl there results quant. the 7- Cl deriv., light yellow, m. $251-2^\circ$; reduction gives nearly quantitatively the 2- NH_2 deriv., light orange, m. $302-3^\circ$. 5,2- $\text{H}_2\text{N}(\text{p}-\text{HO}_2\text{C}_6\text{H}_4\text{CO})\text{C}_6\text{H}_4\text{CO}_2\text{H}$, by reduction of the above NO_2 deriv., gives with oleum 2,7- $\text{H}_2\text{N}(\text{p}-\text{HO}_2\text{C}_6\text{H}_4\text{CO})\text{C}_6\text{H}_4\text{SO}_3\text{H}$, which may be transformed into the 2,7-di- NH_2 deriv. by NH_4OH and As_2O_3 at 180° . Further proof is offered by the m. p. of the corresponding di- HO deriv. and its di- Ac deriv.

C. J. WEST

Some new water-soluble organomercury compounds, JOHN H. WALDO, *J. Am. Chem. Soc.* 53, 992–6 (1931)— Mellig and $\alpha\text{-HSC}_6\text{H}_4\text{CO}_2\text{H}$ in EtOH - NaOH give 80–85% of methylmercurithiosalicylic acid, m. 174° (all m. ps. cor.), *Et* deriv., m. 111° ; *iso-Am* deriv., m. 78° ; *p*-ethylmercurimercaptobenzoic acid does not m. 250° ; the sulfonic acid was also prepd. The following compds were prepd., but analyzed only for Hg : phenylmercurithiosalicylic acid, m. 228.5° (decompn.); benzyl deriv., m. 144.5° (decompn.); α -ethylmercurimercaptobutyric acid, m. 76° ; *p*-ethylmercurimercaptobenzoic acid, m. 116.7° ; *p*-ethylmercurimercaptobenzenesulfonic acid, m. above 300° . These compds are sol. in NaHCO_3 and form with the alkali metals sol. salts whose solns. are in general stable and do not give an immediate ppt. of HgS with $(\text{NH}_4)_2\text{S}$. Therapeutic data are given, the alkaryl derivs. are less toxic than the aryl; the salts of CO_2H derivs. are less toxic than the SO_3H derivs.; some decrease in the therapeutic index is produced by a change from the α - to the β -position on the C_6H_5 ring and by alkyl instead of alkaryl derivs. The aryl compds are decidedly less stable, either in soln. or dry, than the alkaryl compds.

C. J. WEST

Pyromellitic acid, Benzodiketohydrindene and benzodipyridazine derivatives, REINHARD SEKA, HANS SEDLATSCHKE AND HEINRICH PREISSECKER, *Monatsh* 57, 86–96 (1931)— $\text{O}(\text{CO})_2\text{C}_6\text{H}_2(\text{CO})_2\text{O}$ (I) and quinaldine, heated at 250° until the mass is solid, give the condensation product (II), $\text{C}_{12}\text{H}_8\text{O}_4\text{N}_2$, dark reddish brown, decompd. 265° , insol. in EtOH , and the yellow brown, EtOH -sol. compd., $\text{C}_{12}\text{H}_8\text{O}_4\text{N}_2$, decompd. about 125° . Sulfonation of II with 50% oleum at 170° gives a product similar to quinoline yellow, which dyes wool and silk a pure yellow. I and $\alpha\text{-C}_6\text{H}_4\text{NH}_2$ give pyromellitic dinaphthylimide, m. 431° . N_2H_4 , H_2O (4 mols) gives the dihydrazide, light yellow, decompd. about 450° ; its *tetra-Ac* deriv. m. $235-8^\circ$. 1,4,2,5- $\text{B}_2\text{C}_4\text{H}_2(\text{CO}_2\text{H})_4$ and 10 mols N_2H_4 , H_2O in abs. EtOH , heated 24 hrs at 120° , give 90% of 1,6-diphenyl-4,9-dihydroxybenzodipyridazine, carbonizes at 445° ; 1,5,2,4- $\text{B}_2\text{C}_4\text{H}_2$

(CO₂H)₂ and 6 mols. N₂H₄, H₂O give the 1,9-diphenyl-4,5-dihydroxy deriv., decomps. 430° (McCaH₂CO₂C₆H₅(CO₂H)₂ and N₂H₄, H₂O give the 1,9-diphenyl-4,6-dihydroxy deriv., carbonizes at a high temp. 2-Benzoylanthraquinone-3-carboxylic acid and 3.8 mols. N₂H₄, H₂O in abs. EtOH, heated 9 hrs. at 120°, give the hydrazide, light yellow, decomps. 374°. 1,2,4,5-C₆H₃(COCl)₄ and MeONa in C₆H₆ give 88% of *trans* Me succinimide, m. 155°; the *trans*-Ph ester, m. 179.5°, results in 24% yields.

C. J. WEST

The influence of impurities on some physical and crystallographic properties of hemimellitic acid. V. AGARONOV. *Compt. rend.* 192, 99-101 (1931). cf. *Compt. rend.* 124, 555 (1897).—After 2 recrystals of hemimellitic acid, C₆H₃(CO₂H)₃ (I), there can still be isolated small quantities of Ca hemimellitate (II) and another compd (III), both of which are monoclinic and have smaller birefringence than I. The refractive indices for the D line are I, n_x , 1.4316, n_y , 1.6358, n_z (calcd.) 1.6744; II, n_x , 1.52, n_y , 1.68, III, n_x , 1.513, n_y , 1.63. The pure I is triclinic and the axial ratios are 1.705 : 1.06185. The ds. and solubilities at 20° of the 3 compds., and a full description of the crystals of I, are given. Microscopic observation of the process of crystn. shows that small quantities of II and III are included within the crystals of I and are probably the cause of its polychromism in the ultra violet, since according to A., polychromism is always due to the presence of impurities or polymers in regular orientation.

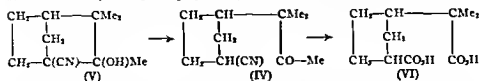
K. V. THIMANN

Sclareol and its derivatives. MAURICE-MARIE JANOT. *Compt. rend.* 191, 847-9 (1930).—Sclareol (I) obtained from the essential oil of *Scleria sclarea* L. appears to be C₂₁H₄₀O₄. It b₁, 188-9° and b₂, 218-22° to give by condensation a solid product, which is transformed slowly at room temp. and instantaneously at 45° into crystals, m. 104-5°. I is not saponified by alc. KOH or by KOH in PhCH₂OH. Heated to near its m. p. it loses 45% of its wt. in 72 hrs. It gives no test for MeO, is acetylated with difficulty, giving only a 50% yield, and is benzoylated in the presence of pyridine with difficulty. It gives CO₂ and CO(NHPh)₂ when heated to 150° with OCNPh. Alc. I gives a violet color with saccharose and furfural, and a garnet red color with PhCH₂ClCH₂ClO, vanillin and piperonal (Elkert, C. A. 22, 2384). In CS₂ contg. Ag₂CO₃ I adds Br to give 3 compds. (1) a cryst. deriv., m. 122-4°, contg. 35.33% Br, and (2) and (3), hexagonal crystals m. 134-5° and 140-3°, resp. I adds I₂ in alc. in the presence of colloidal Pd to give 98% of a cryst. deriv. m. 114-5°, mol. wt. 270. Oxidation with KMnO₄ in HOAc gives sclareolic acid besides other compds. and oxidation with H₂CrO₄ gives a non-acidic compd. m. 124-5°. Alc. I shows no absorption in the ultra-violet. The sp. rotatory data of the various compds. are given.

C. J. PEDERSEN

Camphor and terpenes. VI. So-called camphor cyanohydrins. J. HORRIG AND E. FRANKUCH. *Ann.* 493, 271-304 (1930); cf. C. A. 21, 1908.—II previously showed that camphoramide easily adds HCN, giving a camphor cyanohydrin, Passerini (C. A. 20, 595) claims to have isolated the same compd. from pernitrosocamphor and KCN. A re-examination of P's compd. (I) shows that it is camphene-1-carboxamide. From I is prep'd a Ag salt and a Me ester, C₁₅H₂₂NO₂, m. 102°. The *dl*-, *d*-, and *l* forms of I may be obtained by starting with the proper camphor; thus, the *d* form prep'd. from the pernitrosocamphor from *d*-camphor, HCN and Et₃N in Et₂O has $[\alpha]_D^{25}$ 7.5°. Sapon. of I with alkali gives camphene-1-carboxylic acid, the *d* form has $[\alpha]_D^{25}$ 95.5°. When the *d* and *l* forms are mixed in petroleum ether, the *dl* form seps., showing that the 2 acids differ only in optical properties. O₂ oxidizes the acid to camphenil-carboxylic acid (II) and camphenone acid (III), m. 133-4°. II gives an oxime, m. 171.2°. The action of O₂ on I gives the amide of III. III with aq. alkali gives camphenecarboxylic acid, the *dl* form m. 133-4° and the *d* form m. 142-3°. Reduction of the acid gives dihydrocamphene-1-carboxylic acid, m. 174-5°, $[\alpha]_D^{25}$ -16.7°. Reduction of I gives the corresponding amide, m. 189°. I and Ac₂O, boiled 2 hrs., give *d* camphene-1-carboxylic nitrile (IV), b₁, 108-10°, m. 38-40°, $[\alpha]_D^{25}$ 40.5°. The acid with SOCl₂ and NH₄OH gives I, m. 208°. The camphor cyanohydrin of II, and P. (V) is not saponified by concd. HCl or 48% HBr, with HNO₃ it gives IV and a blue-green oil, which gives with MeOH KOH *d*-camphene-1-carboxylic acid and a small quantity of the amide, there also results an oxynitrile, which is dehydrated by Ac₂O to IV and with EtOH H₂O-KOH gives 3 (α , α -dimethylacetyl)pentamethylenecarboxylic acid (VI), b₁, 180-8°, $[\alpha]_D^{25}$ -2.6°, semicarbazone, m. 200-2°. Further degradation with NaOBr gives camphenecarboxylic acid (VII), m. 134-5°, identical with that obtained from III. The oxynitrile is a camphene hydrate-1-carboxylic nitrile. *d*-Camphene-1-carboxylic acid and HBr in AcOH give the *dl*-Br hydrate, m. 118-20° (decompn.), $[\alpha]_D^{25}$

—1b 5°, dil alkali or K_2CO_3 gives the *camphene hydrate-1-carboxylic acid*, m 142–3°. $[\alpha]_D^{25}$ —17°; the H_2O is split off by boiling with 0.1 N H_2SO_4 for 15 min; the *Me ester*, m 59–60°, $[\alpha]_D^{25}$ —19.5°. $MeOH-KOH$ gives VI. *d*-Camphene-2-carboxylic acid and HI in $AcOH$ give the *Br hydrate*, m 165–6°, the corresponding *hydrate* m 221°. *Me ester*, m 61–2°. $[\alpha]_D^{25}$ 24.5° *dl* Camphene-bromohydrate-1-carboxylic nitrile m 150–1°, attempts to prep the hydrate gave only the unsatd. compd



C. J. WEST

α -Camphor derivatives. II. Identity of dihydroteresantallic acid with 7- α -apocamphanecarboxylic acid. TORSTEN HASSELSTRÖM *J Am Chem Soc* 53, 1097–103 (1931), cf *Ann Acad Fenn* 30, 12 (1929).—The lactone of 2-hydroxy-7- α -apocamphanecarboxylic acid, with alk $KMnO_4$, gives ketodihydroteresantallic acid (2-keto-7- α -apocamphanecarboxylic acid), m 269–70°, $[\alpha]_D^{25}$ —55° (10% EtOH); semicarbazone, m 204–5° (decompn), heating the latter with $EtONa$ at 170–80° for 15 hrs gives a good yield of 7- α -apocamphanecarboxylic acid (dihydroteresantallic acid), m 228–9° (cor), nearly optically inactive, this is identical with the compd obtained by complete synthesis in part I. The crude mixt of Me chlorodihydroteresantalate, reduced with Al Hg in moist Et_2O at room temp, and finally with Na and $EtOH$, gives dihydroteresantalol (α -borneol), m. 171°, oxidation with $K_2Cr_2O_7$ and H_2SO_4 gives dihydroteresantalol, whose semicarbazone m 212–3° (decompn), further oxidation with $KMnO_4$ in $NaOH$ gives dihydroteresantallic acid, m 229–30°, identical with the acid prepd as above. The addn of HCl to teresantalol, m 114°, gives a liquid and a solid form of chlorodihydroteresantalol, the solid part (2-chloro-7- α -borneol) m 125–6° (decompn) and on reduction with Na and $EtOH$ gives α -borneol.

C. J. WEST

Optically active α -pinenes. F. H. THURBER AND R. C. TIDELLS *J. Am. Chem. Soc.* 53, 1030–2 (1931).—Details are given for the prepn. of pinene nitroschlorides, the 3 forms, *d*-, *d*-, and *l*-, m 115°, 89.5° and 90°, resp, and have $[\alpha]_D^{25}$ (Cali) 0°, 390.2° and 366.8°, they decompose at room temp. to a brown oily mass in about 3 weeks, but are stable for a much longer time at —20°, the regenerated α pinenes show the following properties. b_{700} 155–6°, d_{20} 0.8591, n_D^{20} 1.4663, $[\alpha]_D^{25}$ 51.14° and —51.28°. The HCl compds m 132° (cor) and the active forms have $[\alpha]_D^{25}$ (1% EtOH soln) 33.52° and —33.24°. Contrary to previous reports the *dl*-pinene yields an inactive HCl deriv. and the active pinenes yield active HCl derivs. whose rotations are equal and opposite in sign.

C. J. WEST

Stereochemistry of derivatives of biphenyl and analogs. FRANCIS LIONS *J. Am. Chem. Soc.* 53, 1176–9 (1931).—L. outlines certain researches in the 2- and 4-phenylpyridine and 2-phenylquinoline series, phenylpyrroles, phenylindoles and biphenyl derivs. Failure to obtain these substances in enantiomorphous forms has prevented publication; this note is offered because of the work of Adams (*C. A.* 25, 105) in the field of phenylpyridines.

C. J. WEST

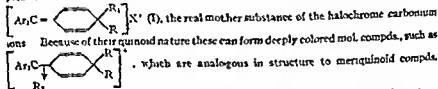
Bistibyls. I. Tetraphenylbistibyl. Attempts to obtain tetraphenylbismethyl. F. F. BLICKS, U. O. OAKDALE AND F. D. SMITH *J Am Chem Soc* 53, 1025–9 (1931).—Diphenylstibyl iodide, m 68–70°, results from the chloride and NaI in Me_2CO ; a $PbBr$ soln. does not absorb O, although a mixt. with Pb_2CBr does. The iodide reacts with Cu in $PbBr$, giving the bistibyl but this apparently undergoes decompn. or is rendered inactive toward O by the Cu halide or by combination with unchanged Pb_2SbI . The iodide in abs. $EtOH$ with NaH_2PO_4 gives tetraphenylbistibyl, m. 121–2°, which reacts instantly with I , giving Pb_2SbI ; it also absorbs the calcd amt of O to give $(PbSb)_2O_2$. Pb_2BiI in $PbBr$, shaken with Ag , Hg , Zn and Cu bronze, did not give solns which absorbed O. NaH_2PO_4 did not appear to react with an $EtOH$ soln at room temp, and on heating gave a black product. $(Pb_2Bi)_2$ could not be obtained.

C. J. WEST

Action of phenol on benzoyldiphenylmethyl bromide. C. FREDERICK KOELSCH *J. Am. Chem. Soc* 53, 1147–50 (1931), cf Schuster, *C. A.* 25, 943.—S reported that Pb_2CBrBz and $PhOH$ give $Pb_2C(OPh)Bz$ (I), which reacted with $PhMgBr$ to give $Pb_2C(OPh)C(OPh)Ph$ (II). Repetition of the reaction shows that the reaction really gives $p-BrC_6H_4OH$ and Pb_2CHBz (S's I). Pb_2CHBz and KOH give Pb_2CH_3 and $BzOK$.

while Ac_2O and H_2SO_4 give $\text{Ph}_3\text{C C}(\text{OAc})\text{Ph}$. Ph_3CHBr and PhMgBr give PhCHCPh OH , m 232-3° (S a II), which decomposes with KOH to Ph_3CH and BzPh . C. J. WEST

Constitution of colorless and colored triphenylmethane derivatives. I. LIRSCHITZ *Ber* 64B, 161-82 (1931), cf *C A* 24, 100, 1368, 4020.—In an earlier paper (Lirschitz and Girbes, *C A* 22, 3855) L. has formulated his theory as to the constitution of derivs of Ph_3CH . In the present paper he answers criticism of his theory by Haantzsch, Dilthey and others, criticizes H.'s theory as to the nature of halochromes, shows the untenability of certain of D.'s views, and develops his own theory further, adding special proofs for certain phases of it. H. has shown that the substances $\text{Ar}_3\text{C X}$ belong to 2 optically different series, between which there is no optical transition. H. considers these series isomeric, if they are, the colorless and colored Ph_3C derivs should have the same mol wts. Since there is no proof of this, L. states that the assumption of isomerism has no foundation. Another defect of H.'s theory is that it cannot explain the most characteristic property of halochrome salts—their peculiar light absorption. In most of the compds $\text{Ar}_3\text{C X}$, the C X bond is heteropolar, $\text{Ar}_3\text{C}^+\text{X}^-$. This may give the electrolyte $[\text{Ar}_3\text{C}^+, \text{R}]^+\text{X}'^-$, whose free ions may be stable or may undergo further change. Anion X'^- may be solvated, or may form complex ions with suitable mols, such as metal halides. The cations may form



Structure I belongs to various so-called halochrome compds with dye like absorption. L.'s theory is thus shown to explain the facts satisfactorily. Retaining a coordination no. of 4 for C as assumed by Werner, and without using the forced deductions of H., it is possible to explain the influence of the solvent on the color of Ph_3CH derivs. The first step in the formation of the colored forms is substitution: the final step is addn with the formation of mol compds of quinoid ions. Where the latter is not possible and complex formation with the compd itself enters in, the reason lies in the nature of the medium. In the substances mentioned the chromophor which shows the visible color is the quinone C atom in the ring. Solns of halochrome derivs of the same triphenylcarbinol contain not the same complex, as H. thought, but different complexes with the same chromophor. With the exception of neutral perchlorates all the isolated halochromes are mol compds, and it can easily be shown that complex formation produces the color. L.'s theory also explains satisfactorily the electrolytic behavior of Ph_3CH derivs, and is further strengthened by the results obtained in detns of the mol wts of such derivs. LOUTER KELLRY

The synthesis of N-methylhydroxynaphthindolequinone and N-methylnaphthaznquinone. ZENJIRO KITASATO AND CHIZABURO SONE *Bull Chem Soc Japan* 5, 348-54 (1930), cf *C A* 22, 1779.—To study the origin and properties of such indigoid substances as nandarurin 1-methylhydroxynaphthaznquinone (VIII) has been synthesized as follows: 2,3-dibromo- α -naphthoquinone and $\text{NaCH}(\text{CO}_2\text{Et})$ give bromo- α -naphthoquinonemalonate ester (I), 5 g. of which suspended in 15 cc. abs. alc. and treated with 5 g. 33% alc. MeNH_2 gives crystals. H_2O dissolves a yellow substance, 1-methyl-3-carbethoxy- α -naphthaznquinone (II), decomps. about 220°, and leaves a red substance, 2-bromo-3-methylacetamido-1,4-naphthoquinone, m 167°. Two g. of II with 120 cc. 7% NaOH gives VIII, m over 300°, 1 g. of which with 10 cc. concd H_2SO_4 and 6-cc. 10 N HNO_3 give 1-methyl-2,3- $\beta\beta$ -naphthaznole-dione-4,9-quinone, m 268°. To elucidate the mechanism of the above reactions the following reactions were conducted: 6,3,4- $\text{Br}(\text{MeO})_2\text{C}_6\text{H}_2\text{CH}_2\text{CHCO}_2\text{Et}$ and 2-bromo-3-carbethoxy-1,4-naphthoquinone both give an amide with MeNH_2 , m 183° and 164-5°, resp. V. F. HARRINGTON

Trivalent nitrogen I. The constitution of indazolecarboxylic acid. TARO HAYASHI *Bull Inst Phys Chem Research (Tokyo)* 9, 911-5 (1930) (Abstracts 88 (in English), published with *Sci Papers Inst Phys Chem Research (Tokyo)* 14, Nos. 275-6).—Crude indazolecarboxylic acid (I) (Auwers *C A* 14, 65) gives, by treating it with 0.5 N KOH , 0.5 N H_2SO_4 , hot AcOEt and PhMe , white crystals m 268-6.5°, whose EtOH soln. is optically inactive. *Brucine salt* m 236-0.5°, $[\alpha]_D^{25} -102^\circ$, crystals from a mixt. of pyridine and EtOH (a) and H_2O (b) show following specific rotations

(a) $[\alpha]_D^{25} -104^\circ$, (b) $[\alpha]_D^{25} -86^\circ$, $[\alpha]_D^{25} -90^\circ$ Canchonine salt, m $233.5-4^\circ$, $[\alpha]_D^{25} 18.1^\circ$, $[\alpha]_D^{25} 17.7^\circ$ Quinine salt, m $245-6^\circ$, $[\alpha]_D^{25} 200^\circ$ The Me ester of I, methylated with MeI, followed by sapon and recrystn from 30% EtOH (K v Auwers, *loc cit*) gives 1-methylindazolecarboxylic acid (II), m $213-4^\circ$ Similarly, on methylation with $(\text{MeO})_2\text{SO}_4$ and purification with MeOH, there results 2-methylindazolecarboxylic acid (III), m $225-6^\circ$ I, II and III show similar absorption curves but the absorption band of I is displaced to the side of shorter wave lengths From these results I would

seem to have the formula $\text{C}_8\text{H}_7\text{N} \begin{array}{c} \text{CCO}_2\text{H} \\ \diagup \quad \diagdown \\ \text{N} \end{array}$ rather than $\text{C}_8\text{H}_7\text{N} \begin{array}{c} \text{CCO}_2\text{H} \\ \diagup \quad \diagdown \\ \text{N} \end{array}$, which contains

an asym C atom

K. KONDA

Condensation of aldehydes and phenols. V. *m*-Chlorobenzaldehyde- β -naphthol. OTTO DISCHENDORFER AND ILFIRICH MANZANO *Monatsh* 57, 20-30 (1931); cf. C A 25, 290— β $\text{C}_{10}\text{H}_7\text{OH}$ (10.2 g) and 5 g $\text{m-ClC}_6\text{H}_4\text{CHO}$ in AcOH-concd HCl, after standing 24 hrs, give ? chloroben-ald β naphthol (I), crystg with 1 mol AcOH, m $178-9^\circ$ Na salt needles *ds Ac deriv* m $216-7^\circ$ *ds Bz deriv* m 164° , Me_2SO_4 gives a mono-Me ether, m 168° I in EtOH and 10% NaOH, treated with Br as long as a yellow ppt forms, gives *drhydro* ? chlorobenzald β naphthol $\text{C}_{21}\text{H}_{13}\text{ClO}_4$, light yellow, m 199° I, AcOH and concd H_2SO_4 , warmed 1 hr on the H_2O bath, give *ms* (1-chlorophenyl)dinaphthopyran or 9 (? chlorophenyl) 1,2,7,8-dibenzoxanthene (II), m $191-2^\circ$, sol in warm concd H_2SO_4 with a golden yellow color and a green fluorescence; II also results in 78% yield by warming the mixt of $\text{C}_{10}\text{H}_7\text{OH}$ and $\text{ClC}_6\text{H}_4\text{CHO}$ in AcOH-HCl 1 hr on the H_2O bath II in AcOH-HCl, oxidized with MnO_2 and treated with solid FeCl_3 , gives *ms* 3-chlorophenylindaphthopyrylium chloride ferrichloride, red needles, m $225-6^\circ$, mercuric chloride, red needles with green metallic luster, m 272° , perbromide, red needles, m $205-6^\circ$ (decompn), perchlorate, bright red needles with green luster The FeCl_3 salt in Me_2CO contg a little H_2O gives *ms* ? chlorophenyl dinaphthopyranol, m $214-5^\circ$, sol in cold concd H_2SO_4 with an orange red color *Et ether*, m $215-6^\circ$, *Me ether*, m $217-8^\circ$, the ethers result from the FeCl_3 salt and the corresponding alc.

C J WEST

The structure of organic molecular compounds with principles of two- and three-dimensional variation. EDUARD HERTEL AND GEORG II RÖMPER *Z physik Chem*, Abt B, 11, 77-80 (1930)—X-ray investigation of org mol compds shows that not all can be built up on the principle of 3 dimensional variation A 2 dimensional variation explains anthracene-1,3,5- $\text{C}_{14}\text{H}_9(\text{NO}_2)_3$ A 1-dimensional variation explains [fluorene]-1,3,5- $\text{C}_{14}\text{H}_9(\text{NO}_2)_3$ The crystal structures of 1,3,5- $\text{C}_{14}\text{H}_9(\text{NO}_2)_3$ and 2,4,6- $\text{MeC}_{14}\text{H}_9(\text{NO}_2)_3$ are discussed

FOSTER D SVETLL

Structure of nitrofuran and the mechanism of nitration in the furan series. B T. FRAYER AND JOHN R JOHNSON *J Am Chem Soc*, 53, 1142-7 (1931)—2 Nitrofuran, prepd by the decarboxylation of 5-nitrofuric acid, is identical with the compd. obtained from furan by nitration, the latter had been incorrectly formulated by Marquis (*Ann chim phys* 4, 216 (1905)) to be 3-nitrofuran Evidence is presented to support the view that the intermediate nitroacetates, produced in the nitration of various furan derivs in Ac_2O soln, are ring structures rather than open chain enols The mechanism of the nitration process is suggested tentatively to be a 1,4 addn to the furan ring, without scission of the O bridge Me furate yields a nitroacetate, $\text{C}_5\text{H}_4\text{NO}_4$, m $96-3^\circ$ (cor), on catalytic hydrogenation 2.5 mols H_2 are absorbed, water decomposes the compd; warming with $\text{C}_4\text{H}_9\text{N}$ gives Me 5-nitrofuroate, m $81-6^\circ$ (cor.)

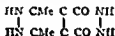
C J WEST

Local anesthetics in the pyrrole series II. F F. BLICKS AND E. S. BLAKE. *J. Am Chem Soc*, 53, 1015-25 (1931), cf C A 24, 1183—In certain types of compds at least, the following substitutions may be made in local anesthetics without loss, in a qual sense, of anesthetic activity the substitution of 2 pyrrol or Bz and 4 amino-benzoyl and the replacement of the NMe₃ and NEt₃ by the 1-pyrrol or the 1-pyrrolidyl nucleus β (1-Pyrrol)ethyl alc. (I), b.p. $110-3^\circ$, and K in C_6H_5 heated until the metal disappeared and then treated with Br_2Cl , gave 78% of the Bz deriv, $\text{BzOCH}_2\text{CH}_2\text{N}(\text{C}_6\text{H}_5)_2$, m $53-5^\circ$. The K deriv of I and $p\text{-O}_2\text{NC}_6\text{H}_4\text{COCl}$ in Et_2O CaH_2 give the 4-nitrobenzoate, slightly yellow, m $92-4^\circ$ catalytic reduction with Pt oxide in abs EtOH gives the 4-aminobenzoate, m $87-8^\circ$ The 2-pyrrolidyl deriv m $73-4^\circ$ $p\text{-H}_2\text{NC}_6\text{H}_4\text{CO}_2\text{CH}_2\text{CH}_2\text{Cl}$ and pyrrolidine, heated for 8 hrs at $115-20^\circ$, give β (1-pyrrolidyl)-ethyl 4-aminobenzoate, m $98-100^\circ$ γ -(1-Pyrrol)propyl alc. b.p. $220-31^\circ$, gives a Bz deriv, b.p. $165-70^\circ$, the 4-nitrobenzoyl deriv m $68-70^\circ$, the 4-aminobenzoyl deriv,

m 114-6°, the 2 pyrrolyl deriv^c m 69-70°. γ (1-Pyrrolidyl)propyl 4-aminobenzoate m 81-5°. The K deriv of MePt(Me₂NCH₂)COH (II) and BzCl give stovaine Methyl ethyl (1-pyrrolidyl)methylcarbamyl benzoate, from the carbunol, BzCl and C₄H₉ on heating 3 hrs on the water bath. The 2 pyrrolyl deriv of II m 104-5°. Dimethyl 1-pyrrolyl methylcarbinol, b₁₀ 86-8°, gives only Bz and 2 pyrrolyl derivs. Diethyl 1-pyrrolyl methyl carbinol, b₁₀ 108-10°, the Bz deriv was only. The corresponding Et deriv b₁₀ 125-8°, the Bz deriv. was only. The 4 (2 pyrrolyl) 1 methyl derivs of the stable and labile forms of 2,2,6 trimethyl-4 hydroxypiperidine were prepd from the 2 forms of 1 methyl vinylidiacetone alkamine (K deriv) and pyrrolyl chloride, both m 106-7° but a mixed m p was 80-3°. Notes are given on the prepn of pyrrolidine. C J West

The preparation of hydroxyproline. HARRIETTE K. KLABUNDE. *J. Biol Chem* 90, 293-5 (1931)—Hydroxyproline is readily obtained as a by-product during the isolation of proline from the products of the hydrolysis of gelatin by 33% H₂SO₄. In Town's procedure (*C. A.* 23, 162) for the sepn of the hydrolysis products proline and hydroxyproline occur in fraction 3, the Cu salts are sol in both H₂O and MeOH. It is sepd from proline by its insol in abs. alc and can be isolated as the picrate from the alk-insol residue. Yields of 18-20 g can be obtained from 1 kg of gelatin. A P LOTIKOFF

Heterocyclic ring systems. III. Ring closure reactions of heterocyclic α -dicarboxylic acids. REINHARD SEKA and HEINRICH PREISSFUCKER. *Monatsh* 57, 71-80 (1931), cf *C. A.* 20, 386—Di Ft 2,5-diphenylfuran-3,4 dicarboxylate (I mol) and 7 mols N₂H₄ H₂O, heated in t OH 24 hrs at 100-20°, give 76% of the cyclohydrazide, C₁₄H₁₀O₂N₂, the corresponding 2,5-di Me ester reacts with the destruction of the ester. Di Ft 2,5-dimethylpyrrole-3,4 dicarboxylate and 4.8 mols. N₂H₄ H₂O, heated 12 hrs at 140-50°, give 80% of the cyclohydrazide, begins to carbonize at 235°, decamps 339°, crystals from H₂O with 1 mol solvent the corresponding 2,5-di Ph deriv m 324° (82% yield). Pyrazolinedicarboxylic dihydrazide, decamps 170°, carbonizes 280° (96% yield), dibenzohydrate, m 218°, the ring system is very labile and could not be isolated. 3 N Phenyl 1,2,3 triazole-4,5-dicarboxylic dihydrazide, m 225° (60% yield), dibenzohydrate, m 240°, heating in vacuo for 2 hrs gives the cyclohydrazide, m 317°. Imidazole-4,5-dicarboxylic acid does not give an anhydride or a cyclic hydrazide. IV. Splitting of the furan ring systems with hydrazine hydrate. *Ibid* 81-5—Di Ft 2,5-dimethylfuran-3,4 dicarboxylate and 2.4 mols N₂H₄ H₂O in abs EtOH, heated 24 hrs at 100-20°, give the compd C₁₁H₁₀O₂N₂, begins to carbonize at 315-20° *tetra Ac deriv*, m 140°, *tetra Bz deriv*, m 190°. The same compd. is obtained by the action of N₂H₄ H₂O upon (LiO₂CClAc). The compd is believed to be 3,6-dimethyldihydropyridazine-4,5-dicarboxylic cyclohydrazide,



C. J. West

Bis(5-phenyl-2-pyrrole)indigo. W. MADELUNG and L. ODERMANN. *Ber.* 63B, 2870-6 (1930)—Posner (*C. A.* 21, 87) considers the indigoid dyes to be intramol. quin hydrones (I), the secondary valences coming into play between the carbonyl O and the aromatic nucleus and the NH serving only to make the aromatic nucleus capable of entering into such a secondary valence union. It might be pointed out, however, that I medlander and Kielbasinski's bis(5-phenyl-2-thiophene)indigo (*C. A.* 7, 2193), in which the conditions demanded by P's theory are not fulfilled, is nevertheless very similar in color and properties to the normal thionaphtheneindigo. It seemed of interest to study more closely the relationship of a pyrroleindigo to the normal indoleindigo and bis(5-phenyl-2-pyrrole)indigo (II) was accordingly synthesized. It seemed that the logical starting material would be a phenylated deriv of 3-hydroxypyrrole (designated as pyrroxy by analogy with indoxyl) which should readily be converted by appropriate oxidizing agents into the corresponding indigo. K phenylglycidate, NH₂CH₂CO₂H and KOH on the water bath yield 80% of the di K salt, needles with 3H₂O from MeOH, m 96°, crystals with 4H₂O, m 86°, from H₂O, of a hydroxy- β (carboxymethylamino)hydrocinnamic acid (III), m 235°. Attempts to apply the Heumann reaction to this acid were only partially successful in that the product of the KOH fusion yielded a very small quantity of an acid (not isolated) which on heating with water gave an oil which must have been the desired pyrroxy for treatment with FeCl₃ and water gave a small quantity of the corresponding II. KOH fusion is not practicable as a method of prepn, however, for by far the greater part of the III is completely destroyed. Attempts to bring about the ring closure by thermal decompa

of the Ca or Th salt of III were likewise only partially successful. Ac_2O with the salt of III, however, yielded 33% of phenylacetylpyrroxylic acid (5-phenyl-1-acetylhydroxypyrrrole 2-carboxylic acid) (IV), m. 150° , decarboxylated by boiling 4 hrs. in water in a CO_2 atm to the pyrrole (38% yield), m. 88° , gradually becoming red in the air. Thus with Ac_2O and a drop of concd. H_2SO_4 gave 33% of 5-phenyl-diacetoxypyrrrole, pale red, m. 181° , also obtained in small yield by heating III w. Ac_2O . With FeCl_3 and HCl IV gave 80% II, brown-violet needles from $\text{C}_6\text{H}_5\text{N}$, somewhat above 300° with evolution of a red vapor, also insol. in most org. solvents in the cold somewhat in hot C_6H_6 and AcOH , easily in hot $\text{C}_6\text{H}_5\text{N}$, the intensely soln. in AcOH is turned pure blue by mineral acids, the H_2SO_4 soln. is orange yell. A 10^{-4} M soln. in $\text{C}_6\text{H}_5\text{N}$ shows extinction below $560\text{m}\mu$, lightening at $480\text{--}500\text{m}\mu$, a 4×10^{-4} soln. in 50% H_2SO_4 lightening at $580\text{--}620\text{m}\mu$, a 2×10^{-4} soln. of the acet. lightening at $510\text{--}560\text{m}\mu$. With complex forming metallic chlorides it forms complexes of different colors (blue with SnCl_4 in C_6H_6 , brown yellow with TiCl_4) w. remain in soln., while dry HCl ppt. a blue HCl salt. The alk. $\text{Na}_2\text{S}_2\text{O}_3$ vat has a slight affinity for cotton and imparts a very weak violet-red color. *Bis*(5-Ac) de. brown red, m. 220° , sol. in AcOH and C_6H_6 , with intense red color, no longer has basic properties or the ability to form metal compds. The properties of II show conclusively that P's theory as to the structure of indigo is incorrect. C. A. I.

Reaction of aliphatic imido ethers with hydrazine. WILFRIED OBERHUNTER, *Monatsh.* 57, 106-11 (1931).— $\text{H}_2\text{N}-\text{CHOEt}$ and anhyd. NH_3 in abs. Et_2O give 25-3 of 1-amino-1,3,4-triazole, besides NH_4Cl and N_2H_4 , HCl , with salts of N_2H_4 in abs. Et_2O only about 3% of the triazole was isolated. C. J. WES.

2-N-Phenyl-5-C-phenyl-1,2,3-triazolecarboxylic acids. MARJO GALLOTTI and G. BARRO and L. SALTO, *Gazz. chim. ital.* 60, 866-72 (1930).—In an analogous way to the transformation of phenanthrene to diphenic acid (I), the oxidation of 2-N-phenyl-1,2-naphtho-1,2,3-triazole leads to 2-N-phenyl-2,3-triazolylphenyl-o,4-dicarboxylic acid (II) (cf. Charrier and Gallotti, C. A. 19, 2265, 2954), but 2-N-phenyl-5-C-phenyl-1,2,3-triazole-4-carboxylic acid, $\text{PhC}_6\text{H}_4\text{N}(\text{NPh})\text{N}(\text{C}_6\text{H}_4\text{CO}_2\text{H})$ (III) (see later) cannot

be prep. in this way because when II is heated alone or with soda lime or with Na it does not simply eliminate a CO_2H group or both CO_2H groups (as with I) with for

tion of III or o- $\text{H}_2\text{C}(\text{CO}_2\text{H})\text{C}_6\text{H}_4\text{N}(\text{NPh})\text{N}(\text{C}_6\text{H}_4\text{CO}_2\text{H})$ (IV). The reaction is much more complicated, with formation of an anhydride deriv., whose constitution is not yet explained. It was of interest, therefore, in explaining the genetic relations between II and III, IV, and their possible transformation into 2-N-phenyl-5-C-phenyl-1,2,3-triazole to attempt to prep. II, III and IV by some other method. The present paper describes the synthesis of III and unsuccessful attempts to prep. II. EtPhCO , treated w. *iso*- AsnNO_2 in the presence of a little anhyd. HCl , extd. with dil. NaOH , the ext. was with Et_2O , AcOH added to the aq. residue and the ppt. recrystd. from dil. Et_2O yields 55-60% of $\text{BzC}(\text{NOH})\text{Me}$ (VI). The latter and PhNHNH_2 yield, after crystn. of the product from EtOH , almost 100% of phenylhydrazone oxime $\text{PhC}(\text{NNHPh})\text{C}(\text{NOH})\text{Me}$ (VII), lustrous, m. 210° . VII (24 g.), added to H_2O (70 g.) at $50\text{--}55^\circ$, after several min. (when clear brown) made ice-cold and when solids cold water (12 vols.) added, agitated for some time, the ppt. washed with cold water and recrystd. from 95% EtOH , yields the *Ac* deriv., $\text{PhC}(\text{NNHPh})\text{C}(\text{NOBz})$ (VIII), lemon yellow, m. 137° . VIII (30 g.) and aq. Na_2CO_3 (14 g. in 950 cc.), heated 40 hrs. on a boiling water bath, cooled, satd. aq. KMnO_4 added until permanently excess KMnO_4 , destroyed by aq. NaHSO_3 and dil. H_2SO_4 , extd. with Et_2O , and the evapd., yields 50% of crude 2-N-phenyl-4-C-methyl-5-C-phenyl-1,2,3-triazole (IX) (15 g.) and alk. KMnO_4 (26 g. and 60 cc. of 10% NaOH in 330 cc. water), refluxed for 100 hrs. (until homogeneous), decolorized with EtOH , filtered, the residue washed with boiling water, the combined filtrates concd., aq. HCl added, the ppt. recrystd. from boiling water, BzOH eliminated by subliming at 180° and the residue recrystd. repeatedly from water, yield III, m. $203\text{--}9^\circ$, stable above its m. p., but at 230° compds. to an orange-yellow substance which sublimes at a still higher temp. decomposition is more extensive than merely the elimination of CO_2 , for in this case would have been formed. III is a strong acid, turns blue litmus red, decomposes (with lively effervescence) and forms alk. salts extremely sol. in water, difficult to obtain in a cryst. state. Prep. like VI, a 45-50% yield of o- $\text{MeC}_6\text{H}_3\text{C}(\text{NOH})\text{Me}$ is obtained, and this with PhNHNH_2 yields almost 100% of the phenylhydrazone oxime, o- $\text{MeC}_6\text{H}_3\text{C}(\text{NNHPh})\text{C}(\text{NOH})\text{Me}$ (X), m. $175\text{--}7^\circ$. Following same procedure as with VII, X (18 g.) and Ac_2O (55 g.) yield almost 100% of the

deriv of X, α -MeC₆H₄C(NNIPh)C(NOAc)Me (XI), light yellow, m. 140°, XI (50 g) and aq Na₂CO₃ (30 g in 1200 cc), heated 40 hrs on a water bath, and treated as before with KMnO₄, yields about 50% of crude 2-N-phenyl-4-methyl-5-tolyl-1,2,3-triazole
C. C. DAVIS

A new synthesis of substituted thioxanthidols. J. REILLY, P. J. DAUM AND B. DALY *Proc Roy Irish Acad* 39B, 615-22(1930); cf C. A. 21, 1253, 2885, 22, 766—(ρ -MeC₆H₄)₂O (I) reacts with C₆H₅(CO)₂O (II) or Br₂Cl (III) and AlCl₃ to give only monosubstituted derivs. H₂SO₄ does not act upon the substituted α -benzoyl-benzoic acid formed with II to yield a substituted anthraquinone but gives 2,7-dimethylfluoran (IV) instead. The product from the reaction of I with III is 9-phenyl-2,7-dimethylxanthidol instead of the expected ketone so the formation of IV probably proceeds through a xanthidol. Exactly similar results are obtained with (ρ -MeC₆H₄)₂S (V) instead of I save that the condensation of V with III proceeds less smoothly. V, b.p. 186° m 57°, was obtained in 70% yield by diazotizing ρ -MeC₆H₄NH₂, destroying the excess HNO₃ with CO(NH₂)₂ and slowly adding the diazonium chloride with vigorous stirring to ρ -MeC₆H₄SH (VI) (1 mol) in dil NaOH (1 mol) at 60-70°, keeping the mixt at this temp for 1 hr after addn is complete. (Explosive diazo thioethers result at lower temps) (ρ -MeC₆H₄)₂S₂ (VII), m 45.6°, was prepd by adding 3.3 g Br₂ to 5 g VI in 100 cc Et₂O at 0°. It dissolves in concd H₂SO₄ to give violet to blue colors, whereas V gives only a faint green, hence it is probable that the violet color obtained by Dillhey and others (C. A. 24, 91, 1935) with V was really due to VII. 2- α -Carboxybenzoyldi- ρ -tolyl sulfide (VIII) was prepd by the Friedel and Crafts reaction and by heating ρ -MeC₆H₄SK with α -(2-5-Cl²MeC₆H₄)CO)C₆H₄CO₂K in the presence of a trace of Cu powder. The Ag salts were sparingly sol in H₂O but sol in dil NH₄OH and the m ps and the mixed m ps were the same, 207°. VIII gave an oxime, m 129°, and warmed with concd H₂SO₄ gave 2,7-dimethylthiofluoran, m 222°. 9-Phenyl-2,7-dimethylthioxanthidol (IX) (from V and III), m 207°. It dissolves in AcOH with a pink color which becomes reddish orange on heating. It dissolves in H₂SO₄ in a reddish orange soln with a brownish yellow fluorescence and the color is not destroyed nor is there ppn on diln with H₂O. Addn of anhyd FeCl₃ to an HCl soln of IX in anhyd Et₂O with HCl gives deep red prisms of the chloride HCl of IX which is readily sol in CHCl₃ and in PhNO₂ but insol in Et₂O and in PhH. This compd. probably has a quinonoid rather than a sulfonium structure (cf Ber 37, 2937; C. A. 4, 1172). IX is slowly reduced by boiling in alc HCl to the corresponding thioxanthene, m 157°, but reduction by Zn in 80% AcOH is much more rapid. PhNEt₃ reacts with IX to give 4'-diethylamino-9,9-diphenyl-2,7-dimethylthioxanthene, colorless needles, m 211°.

C. H. PEET

Purely aliphatic strepto-pentamethine dyes. W. KÖNIG AND W. REGNER. Ber. 63B, 2823-7(1930)—That compds. of the type Alk₃N(CH)₂-1'-NAlk₃X or, rather, [Alk₃N—CH—CH—CH—CH—CH—NR₂]_n (I, R₂N = Me₂N and CH₃ CH₂ CH₂ CH₂ CH₂ N)

sufficient color strength and fiber affinity for practical dyeing purposes had not as yet been definitely proved, although indications of this had been obtained with the pentamethinium salts ($n = 3$). Zincke, at the beginning of his classical work on the cleavage of pyridine, had attempted to prep the salts

but obtained them only in the form of very difficultly sol HgCl₂ addn compds, and he distinctly expressed the opinion that the salts themselves are so unstable that they cannot be isolated as such, and as the yellow color of the HgCl₂ compds. might to a certain degree be produced by the HgCl₂ the question whether these salts are real dyes remained unanswered. The desired evidence has now been supplied by the synthesis of the well-crystd perchlorates (I, X = ClO₄). By the BrCN C₆H₅N method they are formed only to a minor extent, together with many other products, which renders their purification difficult, they can also be obtained from NMe₃ or piperidine with various di-N-arylamino-pentamethinium salts but these methods are far inferior to that depending on the treatment of the Zincke aldehydes, ArAlkNCH=CHCH=CHCHO, with NMe₃ or piperidine in the presence of 1 equiv of HClO₄. The 2 perchlorates are, in solid form, yellow with beautiful blue surface luster, they are relatively stable, easily dissolve in H₂O and alc with green yellow color and dye tannated cotton similarly to the auramines as regards tone and depth of color. The absorption curves of the 2 salts are almost identical. The NMe₃ compd shows a single main band with a max at 412 μ which is materially more persistent and narrower than the corresponding

band complex of auramine perchlorate; at around 300μ the dye is extraordinarily transparent. Colorimetric comparison of ale solns of the Me_2N compd and com. auramine G showed relative strengths of about 85/100. The fastness to light of the color on tannated wool was at least equal to that of auramine and the fastness to acids and alkalis was greater. A further striking property of the new salts is their strong luminescence under the Hanauer quartz lamp (intense green and orange-yellow for the Me_2N and piperidine compds, resp.), this fluorescence is shown only by the solids, not their solns, and is limited to the perchlorates. The 1,5-bisdimethylaminopentamethinium perchlorate m 165° , the 1,5-bis-*N*-piperidino salt m 111° . C. A. R.

Linear pentacene series. XX. Optical absorption of pentacene compounds. GUIDO MACHEK *Monaish* 57, 201-24(1931), cf. C. A. 24, 5019.—Details are given, in tables and as curves, of the absorption of pentacene 6,13 monoquinone, pentacene-5,7,12,14 diquinone, 1,8-dihydroxypentacene-5,7,12,14-diquinone, the 6,13 di HIO deriv., the 1,8-di-Br deriv., the 6,13 di Br deriv. and the 1,11-di- NO_2 deriv. in C_{11}H_8 and $\text{C}_{11}\text{H}_9\text{N}$ and of the 1,11 di Br deriv., the 2,9 di-Br deriv., the 1,8- and 1,11-diamino derivs., and the 1,8-di NO_2 deriv. in $\text{C}_{11}\text{H}_9\text{N}$. The influence of the solvent and of the substituents of the absorption spectra is discussed. C. J. WEST

Alkali-organic compounds. VIII. Reactions between lithium alkyls, pyridines and condensed pyridine systems. K. ZIFGLES AND H. ZEISER *Ann.* 485, 174-92 (1931), cf. C. A. 24, 5752.—Quinoline and a 1,32 *N* BuLi soln in C_6H_6 give an addn. product (I) which, on decompn. with H_2O , gives 90% of 2-butyl-1,2-dihydroquinoline (II), b_p 128°, b_{11} 100-2°, d_4^{20} 0.907, n_D^{20} 1.5699, n_D^{25} 1.5772, picrate, deep red, m $140-60^\circ$. Thermal decompn. of I gives 60-60% of 2-butylquinoline (III), b_{11} 153° after purification through the picrate, m $162-3^\circ$, the crude III also yields a picrate, m 145° , probably that of the 4-Bu deriv. III has d_4^{20} 1.003, n_D^{20} 1.5699, n_D^{25} 1.5768. Heating II in 5 parts PhNO_2 gives 70-80% III. PhLi gives principally 2-phenylquinoline and a little 4-Ph deriv. Isoquinoline and BuLi give an unstable dihydro-1-butylisoquinoline, which b_p about 135° and on dehydration by boiling in PhNO_2 gives 70% of 1-butylisoquinoline, b_{11} 154-7°, b 283-4°; picrate, m 185.5° . PhLi gives the 1-Ph deriv. Acridine and BuLi give 0-butyl 9,10-dihydroacridine. 2-Butylpyridine and BuLi give dibutylpyridine, b 243-4°; chloroplatinate, yellow, m $102.5-3^\circ$. Quinaldine gives a Li deriv., which yields with PrBr III; PhCH_2Cl gives 2-(β -phenylethyl)quinoline, m 28° . Allyl chloride gives 2-(3-butenyl)quinoline, b_{11} 152.5-4°, d_4^{20} 1.023, n_D^{20} 1.5853, n_D^{25} 1.5920, picrate, yellow, m 143° . Ph_2CO gives 1,1-diphenyl 2-(2-quinolyl)ethanol, m 165° ; H_2SO_4 in AcOH splits off H_2O and gives 2-(β , β -diphenylvinyl)quinoline, m $103-3.5^\circ$. The Li compd. of picoline and PhCH_2Cl give 2-(β -phenylethyl)pyridine. C. J. WEST

Ultra-violet absorption spectra of the quinoline group. C. STANTON HICKS. *Australian J. Exptl. Biol. Med. Sci.* 7, 171-81(1930).—Cinchonine, quinine, cinchonine, hydrocupreine, quinoline, cupreine, methylquinine, ethylquinine, ethylhydrocupreine, eucupine, yohimbine and equisinine were studied. The quinoline nucleus is the detg. factor in the absorption characteristics of the group. The absorption curves of stereoisomers were practically identical. Substitution in the quinoline HIO group, reduction of the vinyl group, esterification, or opening of the quinuclidine nucleus, exert little or no effect upon the curves. The absorption curve of yohimbine is similar to that of indole and quinoline. C. G. KING

The alkyl-py-quinolines. (Generalization of the Skraup's reaction applied to α -alkylglycerols.) RAYMOND DELABY AND MILLE JEANNE HIRON. *Compt. rend.* 191, 845-7, *Bull. soc. chim.* [4], 47, 1395-400(1930).—The alkyl-py-quinolines were obtained in low yields by means of the Skraup's reaction modified by the authors. They are clear yellow liquids possessing a strong odor, and were found to be mixts. of α - and γ -isomers, the former being present in greater quantities. The phys. properties of the following quinolines and their derivs. are: α -ethyl-, α -propyl and α -butylquinoline, resp. b_{11} 128-31°, 142-5°, 150-5°, d_{17} 1.050, 1.038, 1.020 (at 18°), n_D^{25} 1.5979, 1.5886, 1.5709, picrate, m 148° , 150° , 143° , chloroplatinate, m 188° , oily, 163.5° ; chloromercurate, m 118° , 112° , —, iodomercurate, m 135° , —, —, iodobismuthate, m 205° (decompn.), 180° , 163° . C. J. PEDERSEN

Methoxylated phenylquinolines or 2-phenyl-4-hydroxyquinoline. REINHARD SEKA AND WALTER FUCHS *Monaish* 57, 52-62(1931).—These compds. were prepd. as a preliminary step to the synthesis of certain flavones. PhNHIBz gives 68.5% of a chloroimide, b_p 200-10°, m. 40° , which condenses with $\text{CHNa}(\text{CO}_2\text{Et})_2$ in Et_2O , by heating 16 hrs. at $130-50^\circ$, giving 27.6% of the Et ester, m $261-2^\circ$, of 2-phenyl-4-ketoquinoline-

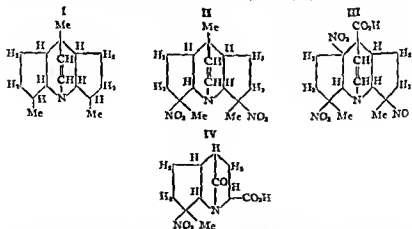
carboxylic acid, which reacts with PCl_5 and NH_3 to give the *amide*, *m* 209°, and with $\text{N}_2\text{H}_4 \cdot \text{H}_2\text{O}$ to give a *pyrazolone compd.* $\text{C}_{11}\text{H}_{11}\text{ON}_2$, *m* 317°. *p*- $\text{MeOC}_6\text{H}_4\text{NHNHBr}$, *m* 156.5°, gives a chloroimide, *b_m* 220°, *m* 63-4°; $\text{ClNa}(\text{CO}_2\text{F})_2$ gives 20.5% of the *Et ester*, *m* 215°, of 2-phenyl-4-keto-6-methoxyquinoline-3-carboxylic acid, *m* 235°; sublimation at 12 mm splits off CO_2 , giving 2-phenyl-4-keto-6-methoxyquinoline, *m* 287°; *p*- $\text{MeOC}_6\text{H}_4\text{NHNHCO}_2\text{C}_6\text{H}_4\text{OMe}$, *m* 202°, yields a chlorosulfide, *b_m* 215-25°, *m* 109°; condensation with $\text{ClNa}(\text{CO}_2\text{Et})_2$ gives the *Et ester*, *m* 205-6°, of 2-*p*-methoxyphenyl-4-keto-6-methoxyquinoline-3-carboxylic acid, *m* 252°, the AcOH soln shows a pale bluish yellow fluorescence. Sublimation in a high vacuum gives 70% of 2-*p*-methoxyphenyl-4-keto-6-methoxyquinoline, *crystg* with 1 mol H_2O , *m* 295°.

C. J. West

Pyrimidines. CXX. Action of ozone on uracil. TREAT B. JOHNSON AND ROBERT B. FLINT *J. Am. Chem. Soc.* 53, 1077-81 (1931), cf *C. A.* 24, 2462.—Uracil is not readily attacked by O_3 in glacial AcOH or 85% HCO_2H at room temp.; if, however, the pyrimidine is suspended in either solvent and washed O_3 is bubbled into such a suspension for several hrs., the uracil is finally completely destroyed, the 2 major products are oxaluric acid and formylglyoxalurea, $\text{C}_4\text{H}_4\text{O}_6\text{N}_2$, *m* 162°, which does not yield a phenylhydrazone, the *Ag salt* crystallizes with 1 mol H_2O , it reduces Fehling soln and Tollens reagent, dil H_2SO_4 gives HClHO and $\text{CO}(\text{NH}_2)_2$. Part of the uracil mol is completely broken down with the formation of $\text{CO}(\text{NH}_2)_2$, $(\text{CO}_2\text{H})_2$, and HCO_2H . **CXXI.** Action of ozone on some derivatives of uracil. *Ibid.* 1082-7.—4-Methyluracil and O_3 give 67% of acetylglyoxyurea, *m* 150°, $\text{CO}(\text{NH}_2)_2$, and $(\text{CO}_2\text{H})_2$; 4-phenyluracil gives benzoylglyoxyurea. Thymine gives 40% of $\text{H}_2\text{NCONHClHO}$, AcCO_2H was not detected. 5-Bromouracil gives parabanic acid. 5-Nitrouracil reacts slowly (28 hrs.), giving parabanic and oxaluric acids. 1,3-Dimethyl-5-bromouracil gives dimethylglyoxyurea, *m* 124-5°. These results show that the pyrimidine ring is destroyed with the formation of characteristic ureide constructions and simple oxidation products. In no case is the urea configuration of the respective pyrimidines destroyed by the O_3 treatment. The application of ozonization makes possible an improved technique for detg. the structure of uracil compds. The advantages gained exactly become apparent when one is called upon to sep. and identify the products of oxidation.

C. J. West

Nitrogen compounds in petroleum distillates. III. Structure of a hydroaromatic base of the formula $\text{C}_{14}\text{H}_{15}\text{N}$. W. C. THOMSON AND J. R. BAILEY, *J. Am. Chem. Soc.* 53, 1002-11 (1931), cf *C. A.* 24, 1804.—The base (I), after purification through the picrate, *b_m* 278.2°, *d₄²⁰* 0.9391, *n_D²⁰* 1.5129, optically inactive, *picrate*, yellow, *m* 151°, *acid sulfate*, *m* 106°, *HCl salt*, *m* 251°, *nitrals*, crystals with 1 mol H_2O .

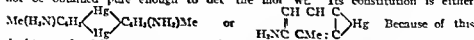


m 79°, is dehydrated at 110° and then *m* 141°, soly. in H_2O at 30° 1:20, *chloroplatinate*, *decomps* 240°, *zinc chloride salt*, *m* 171°; *mercurichloride*, *soltns* at 250° and then *decomps* without melting. I does not contain a *N* alkyl group and cannot be dehydrogenated by $\text{Hg}(\text{OAc})_2$ in 50% AcOH or by other methods, *Br* gives only a

loose Br addn product. Oxidation with KMnO_4 gives only $(\text{CO}_2\text{H})_2$; I is not reduced by Na and abs. $\text{Fe}(\text{OH})_3$, Sn and HCl or HI and red P. I is regarded as *decahydro-3,5,8-trimethyl-4,8-ethenopyrrolidine*. Nitration of I gives a mixt. of 4 products when heated at 170° for 3 hrs., 3,5-di NO_2 deriv. (II), m 116.5° (picrate, m 231°), *decahydro-3,5-dimethyl-3,5,7a-trinitro-4,8-ethenopyrrolidine-8-carboxylic acid* (III), yellow, m 347° (decompn), *octahydro-7-methyl-7-nitro-1,4-carbensilpyrrolidine-3-carboxylic acid* (IV), m $189-90^\circ$ (decompn), readily oxidized with alk. KMnO_4 , berberonic acid, m $240-1^\circ$ (decompn). C. J. WEST

Oxidation of heterocyclic arsenic compounds by iodine. G. A. RAZUYAEV AND V. S. MALINOVSKI. *Ber* 64B, 120-30 (1931), cf. C. A. 24, 2132. — A new method for the detn. of As in alkyl and aryl derivs. of 9,10-dihydrophenarsazine (I), in 10-chloro-9,10-dihydrophenarsazine (II), in 9,10-dihydrophenarsazine oxide (III) and in phenoxarsine oxide (IV). A soln. of the sample in 15-20 cc. EtOH was dild. with H_2O until cloudy, and titrated with a 1 (0.1 N for 0.3-0.5 g. samples, 0.02 N for 0.1 g. sample) to a bright yellow color. During titration H_2O was added 2 or 3 times so that the total vol. was 60-100 cc. The color appearing on the addn. of the 1st portions of I disap. appeared at once on shaking. Toward the end of the titration, 10-30 sec. was required for disappearance of the color. Finally 1-2 drops of I soln. gave a bright yellow color which lasted 24-48 hrs. This method is good for all derivs. of I which are not colored or contaminated with colored substances. The oxidation products must also be colorless or nearly so. The method is not good for I, since a brown color developed and obscured the end point. With IV, which is insol. in EtOH , the method was modified. With both IV and II, the addn. of NaHCO_3 during the titration was advantageous. The results obtained by this new method were within the permissible limits of error. Before the oxidation products could be studied it was necessary to remove the III also formed. This was done (1) by neutralization with NaOH , or (2) by pptn. with an excess of freshly pptd. Ag_2O and removal of the ppt. of AgI and Ag_2O by filtration. The products were obtained on concn. of the soln. and with a nearly theoretical yield in each case. Method 1 gave mainly the di-OH derivs. of the 10-R derivs. of I; method 2 gave the corresponding oxides. The products usually contained some H_2O of crvstn. or hygroscopic moisture, which was removed by drying *in vacuo* (5-10 mm.) over P_2O_5 . The di-OH derivs. dried *in vacuo* at $70-150^\circ$ gave the oxides. The following oxidation products were isolated: di-OH deriv. of 10-Me deriv. of I, m. $210-15^\circ$ (80% yield), the corresponding oxide (also obtained by drying the di-OH deriv. over P_2O_5) decomp. without melting at $236-7^\circ$, the di-OH deriv. of the 10-Et deriv. of I m. $141-3^\circ$ when contg. H_2O , it m. $141-153^\circ$ when H_2O -free; the corresponding oxide m. 230° , the di-OH deriv. of the 10-Pr deriv. of I m. 93° , the corresponding oxide m. $111-2^\circ$, the di-OH deriv. of 10-iso-Am deriv. of I m. $95-6^\circ$, the corresponding oxide m. $101-2^\circ$, hygroscopic; the 10-Ph deriv. of I was not obtained, the product being always the oxide, m. $280-300^\circ$ (decompn). Oxidation of III gave phenarsazinic acid (V); of II gave V; of IV gave phenoxarsinic acid, m. $221-2^\circ$; of AsPh_3 gave triphenylarsine dihydroxide, m. $112-4^\circ$. These oxidation products gave dark red or cherry-red solns. in concd. H_2SO_4 . The following derivs. of I were prepd.: 10-Pr deriv., from MgPrI and II, m. $81-2^\circ$, 10-iso-Am deriv., from iso-Am MgBr and II, m. $76-8^\circ$; 10-Ph deriv., from PhMgBr and II, m. $148-9^\circ$. LOUISE KELLEY

The formation of a closed heterocyclic ring of mercury atoms. II. L. VECCHIOTTI AND M. PANCIERA. *Gazz. chim. ital.* 60, 904-9 (1930); cf. C. A. 25, 957. — The constitution of the Hg ring compds. prepd. in the previous expts. (cf. V, C. A. 23, 3447) by the action of $\text{Na}_2\text{S}_2\text{O}_3$ on PhNH_2 derivs. contg. 2 nuclear AcOHg groups was not definitely established, and in view of this and because of their interesting character, the expts. were continued with the prepn. of other compds. 2,3,4-Me(AcOHg) $\text{C}_6\text{H}_3\text{NH}$ -3,4-Diacetomercuro-*o*-toluidine (cf. V. and Copertini, C. A. 24, 348) (6 g.) and aq. $\text{Na}_2\text{S}_2\text{O}_3$ (10 g. in 10 cc.), let stand 1 month in darkness, filtered, the residue washed with water and dried over H_2SO_4 , yield 3,4-*mercuro-bi-o-toluidine* (I), m. 200° , could not be obtained pure enough to det. the mol. wt. Its constitution is either



Because of this doubt, a quite different method was tried to obtain a compd., the identity of which could be more easily established. *o*- $\text{C}_6\text{H}_4\text{Br}_2$ (50 g.), 2% Hg-Na (500 g.), dry gasoline (50 g.) and AcOEt (5 g.), refluxed 5 hrs. at 120° , filtered hot, the gray product in the residue sepd. mechanically from the Hg , dried in air, washed with boiling water, dried recrystd. from boiling cumene (because of the insoly. in all ordinary solvents) and washed with EtOH and then with Et_2O , and dried *in vacuo* over H_2SO_4 , yields 1,2-di-

mercurisbiphenyl (II), turns light brown when heated, but does not m up to 300°. Its analysis corresponds to $C_{12}H_{10}Hg$, so that it is either $C_6H_5\begin{smallmatrix} \diagup Hg \\ \diagdown Hg \end{smallmatrix}C_6H_5$ or $C_{12}H_{10}Hg$

To decide between these 2 formulas, its decompn was studied. II and glacial AcOH, refluxed for a long time, filtered hot, the filtrate cooled, and fractionally crystd and the product dried in vacuo over H_2SO_4 yield a *bis*(acetomercurs)benzene, $(AcOHg)_2C_{12}H_{10}$ (III), does not m up to 300°. The mother liquor from this crystn, neutralized with KOH and steam distd, the distillate extd with H_2O and the Pt_2O evapd from the ext yields $C_{12}H_{10}$. This is in accordance with the 1st formula of II given above, and with the

reaction $C_6H_5\begin{smallmatrix} \diagup Hg \\ \diagdown Hg \end{smallmatrix}C_6H_5 + 2AcOH \rightarrow III + C_{12}H_{10}$. III (6 g) and I (7.68 g)

in aq KI (6 g in 60 cc), intimately mixed, let stand 24 hrs at room temp, steam-distd, the distillate extd with Pt_2O , the ext evapd, the residue cooled with ice, the solidified product dissolved in a min of boiling EtOH, the soln made ice-cold and the ppt dried in vacuo over H_2SO_4 yields α - $C_{12}H_{10}$, m 27° (el Körner and Wender, *Gazz chim ital* 17, 491). This proves conclusively that II has the 1st formula given above. C C DAVIS

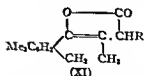
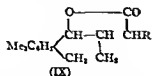
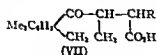
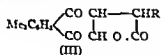
Attempted synthesis of a tricyclic system present in morphine. RICHARD H MANSKE *J Am Chem Soc* 53, 1104-11 (1931).—An attempt, thus far unsuccessful, to effect the synthesis of a portion of the morphine nucleus, is reported. Details are given of the prepn of phenylsuccinic and β phenylglutamic acids. Reduction of $EtO_2C\cdot ClPhCH_2CO_2Et$ with Na and EtOH gives 2 phenylbutane-1,4-diol, b_p 165°, sets to a solid glass at -10°, *bis*phenylurethan, m 113°, dry HBr in Ac_2O gives the dibromide, b_p 173-5°, loses HBr readily and in several days turns turbid and dark, NaCN, followed by hydrolysis, gives β phenylglutamic acid, m 146°, with PCl_5 and then $AlCl_3$ in CS_2 , there results 1-keto-1,2,3,4 tetrahydronaphthalene-4-acetic acid, but the yield was so small that it was not obtained in a state of purity, there also resulted a small quantity of 4,9-diketo 1,2,3,4 tetrahydroacenaphthene, m 149°. 3-Phenylpentane 1,5-diol, b_p 174°, *bis*phenylurethan m 112-3°, the dibromide, b_p 177-82°, loses HBr readily and with EtOH $PhONa$ gives 1 phenyl 1,5-diphenoxypentane, m 72°. γ -Phenylpicemic acid, m 84°, PCl_5 , followed by $AlCl_3$, gives 53% of 1-keto 1,2,3,4 tetrahydronaphthalene 4 propionic acid, m 108-9°, the Me ester is a viscous oil, there also results in this condensation hexahydrobenzacenaphthene α,α' -dione, m 111°; dioxime, m 253-4°. Phenylglutamic acid, PCl_5 and $AlCl_3$ give hydrazid-1-one-3-acetic acid, m 155°. Me ester, b_p 175-80°, m 54°. In the reduction of esters with Na, the use of EtOH is as satisfactory as the toluene method, in the case of the ester contg $PhO(CH_2)_3CN$, there results α phenoxybutyl 3-phenoxybutylamide, m 94-5°. $PhO(CH_2)_3OH$ is obtained in 75% yield with the Na-EtOH reduction, $HIO(CH_2)_3OH$ is obtained in 77% yield. C J WEST

Syntheses in the field of the santonin derivatives. A E CHICHIBABIN AND M N SHCHUKINA *Ber* 63B, 2793-806 (1930).—This paper is published because of the appearance of the communications of Clemon, Haworth and Walton (*C A* 24, 4046) and of Berg (*C A* 24, 4300). The method chosen for the synthesis of santonin and related substances is closely analogous to that used for the synthesis of pilonic acid (*C A* 24, 3016). Succinic esters and their homologs are condensed with oxalic esters to oxalosuccinic esters $R'O_2CCH(R)CH(CO_2R')CH(OH)CO_2R'$, which are reduced to the HO esters, $R'O_2CCH(R)CH(CO_2R')CH(OH)CO_2R'$, and these on heating yield the lactonic acid esters $O \text{ CO CHR CH(CO}_2R') \text{ CHCO}_2R'$ which contain 3 asym C

atoms and may therefore exist in 4 optically inactive stereoisomeric forms. By condensation of the anhydrides or chloroanhydrides of these lactonic acids with aromatic compds (especially *p*-xylene and 2,5-Me₂-C₆H₃OMe) it was planned to prep the products of incomplete condensation, 2,5-Me₂-C₆H₃COR' (I) and 2,5,4 Me₂(MeO)C₆H₃COR' (II) ($R' = -CH(O)COCHRCHCO_2H$), and those of complete condensation (III

and IV) (IV = III with Me₂(MeO)C₆H₃ instead of Me₂C₆H₃). I and II on reduction should yield the dibasic acids Me₂C₆H₃CH₂CH₂CH(CO₂H)CHRCO₂H (V) and Me₂(MeO)C₆H₃CH₂CH₂CH(CO₂H)CHRCO₂H (VI) which by ring closure should give the tetrahydronaphthalene derivs VII and VIII. Reduction of the ketone C O group in VII and VIII and subsequent lactone formation should give compds with structures (IX and X) which, when R = Me, have been shown by the work of C. H and W to be the structures of hyposantonin and desmotroposantonin Me ether. IX and X

are also obtained by the method successfully used by C., H. and W., viz., ring closure of VII and VIII to the unsatd lactones, XI and XII, and reduction. Reduction of the two ketone C O groups in III and IV should give compds (XIII and XIV) having the structures which recently have usually been assigned to hyposantonin and desmotroposantonin Me ether. Sapon of the MeO group in X and XIV would give the compds (XV and XVI) having the 2 structures which have been given to desmotroposantonin. XIII and XIV can also be obtained in a more round-about way—reduction of I and II to $\text{Me}_2\text{C}_4\text{H}_7\text{CH}_2\text{R}^*$ (XVII) and $\text{Me}_2(\text{MeO})\text{C}_4\text{H}_7\text{CH}_2\text{R}^*$ (XVIII) (R in $\text{R}^* = \text{Me}$), ring closure to the tetrahydronaphthalene derivatives (XIX and XX) and the reduction of the C O groups to CH_2 . All these syntheses might be rendered difficult by the presence of the 3 asym C atoms but it was hoped that the intermediate products corresponding to the stable hyposantonin and desmotroposantonin might also be stable and therefore the most easily formed modifications. The work has not yet been completed and the results so far obtained are published now to reserve the right of further investigation along this new broad road to the synthesis of santonin-like lactones. From $\text{EtO}_2\text{C}-\text{CHMeCH}(\text{CO}_2\text{Et})\text{CO}_2\text{Et}$, $(\text{CO}_2\text{Et})_2$, and NaOEt was obtained 80-90% $\text{EtO}_2\text{C}-\text{CHMeCH}(\text{CO}_2\text{Et})\text{CO}_2\text{Et}$ (probably a mixt of stereoisomers), converted by reduction with amalgamated Al in moist Et_2O into 85% of a mixt. of esters of 3 stereoisomeric HO acids. On distn *in vacuo* elimination of EtOH and lactone formation occurred and 200 g of the reduction product after 15 fractionations yielded 88 g liquid ester b_p 182-3°, d_{20} 1.1717, n_D 1.4498, 18 g b_p 186-7°, d_{20} 1.1747, n_D^{20} 1.4507 and 13 g solid ester, b_p 200-4°, m 70°. Of these *di Et 3 methylbutanolide-1,2-dicarboxylates*, the first yielded with boiling HCl a free acid m 170-82°, which with AcCl gave, along with dimethylmaleic anhydride, the anhydride, m 162°, aniside, from the anhydride and boiling PhNH_2 , m 212-4°. The ester b_p 186-7° gave an acid m 186°, which strongly depressed the m, p of the preceding acid and was for the most part unchanged by boiling AcCl , the small quantity which did react giving the above anhydride. The 2 acids are apparently *cis trans* isomers. The solid ester gave an acid m 185°, forming with AcCl an anhydride m 201° which partly isomerizes into the 162° anhydride on distn *in vacuo* and yields the same aniside with PhNH_2 . The 162° anhydride with *p*-xylene and AlCl_3 gave the acid I, m 171-3°, converted by heating 5 min in concd. H_2SO_4 on the H_2O bath into an isomer m 150°, which is obtained directly from the chloroanhydride of the 181-2° acid with *p*-xylene and AlCl_3 . I and its isomer with amalgamated Zn and concd HCl gave the compd V, m 161-3°, but with amalgamated Zn and AcOH they yielded a compd 2,5- $\text{Me}_2\text{C}_4\text{H}_7\text{COCH}_2\text{CH}(\text{CO}_2\text{H})\text{CHMeCO}_2\text{H}$, m 169-72°, which is reduced to V by the Clemmensen method. The chloroanhydride, m 182° (decompn), of I, heated at 160-85°, yields the diketone III, light yellow, m 137-9°, reacts neutral to litmus, sol in boiling 10% but only difficultly in 0.1 N NaOH . The 162° anhydride with 2,5- $\text{Me}_2\text{C}_4\text{H}_7\text{OMe}$ gives the ketolactonic acid II, m 207-8° (together with a small quantity of a substance m 156°), which is reduced by the Clemmensen method to the compd. VI, m 131°, converted by demethylation with HI into the cryst product from which was obtained, by the method of C., H. and W., their unsatd lactone m 250-2°.



C. A. R.

Sulfur sugars and their derivatives. XV. Nature of the sugar of mustard oil glucosides. WILHELM SCHNEIDER, HELLMUTH FISCHER AND WALTER SPECHT. *Ber.* 63B, 2787-98(1930); cf. *C. A.* 23, 4032.—It has been established that the natural mustard oil glucosides are *S* glucosides of the type $\text{RN C(OSO}_2\text{M)SC}_4\text{H}_9\text{O}_4$ ($\text{M} = \text{K}$, an org. base, etc.) and it only remained to det. whether they are α - or β glucosides. From their negative rotation and the fact that all plant glucosides hitherto investigated have been found to be derivs. of β glucose, it was to be expected that these *S* glucosides would also belong to the β series, although Wrede had reported that the thio-

sugar obtained from sinigrin is not identical with the thioglucose synthesized by Schneider and collaborators which, from its method of prepn. and reactions, seemed to be beyond doubt a β -compd. The subsequent prepn. in pure cryst. form of the Na salts of the 2 mutamer thioglucoses and the exact characterization of the optical behavior of the free thiosugar undermined the ground on which Wrede based his conclusion that the thiosugar from sinigrin is the α form and again left open the question of its configuration. Since the 2 typical glucosidases, emulsin and yeast maltase, can not be used, the mustard oil glucosides being hydrolyzed only by myrosinase, which is sp. for them alone, and since the synthesis of these glucosides has not yet been effected, their configuration could be settled only by detg. that of the glucose set free by their cleavage. Preliminary expts. along these lines on sinigrin have already been described (S. and Becker, *C. A.* 24, 2466). Contrary to this preliminary announcement, however, the initial rotation of the glucose liberated by myrosinase is not unequivocal. It was found that it is impossible to prep. an optically inactive myrosinase soln. and its rotation is so affected by NiCl_2 as to render it impossible to recognize with certainty any mutarotation effect of the sugar split off. The expts. described in the present paper, however, confirm the results given in the preliminary note and lead to the same conclusion on a firmer exptl. basis. The mustard oil glucosides are smoothly and rapidly decompd. by aq. AgNO_3 with pptn. of allyl mustard oil AgSO_4 in the case of sinigrin and a mixt. of the mustard oil and sinapic acid Ag compds. in that of sinalbin, and if the HNO_3 set free is at once neutralized by carrying out the reaction in the presence of Ba or Ag carbonate, the reaction is practically complete in 0.5 hr. and the mutarotation of the liberated glucose is not unduly accelerated by the acid. Hg^{++} salts can also be used, although not so satisfactorily, for the decompn. From sinigrin, m. 127-0°, $[\alpha]_D^{20} -16.4^\circ$ (H_2O , c 2.622), with 3 mols. AgNO_3 was obtained a glucose soln. with $[\alpha]_D^{20}$ changing from 81.0° 40 min. after adding the AgNO_3 to 64.7° after 120 min. and 53.7° for $t = \infty$ (20-30 hrs.). Sinalbin, m. 84-5°, $[\alpha]_D^{20} -8.76^\circ$ (H_2O , c 0.2967), gave a glucose soln. with $[\alpha]_D^{20}$ 93.7°, 53.4° (const.) after 40 and 120 min., resp. With $\text{Hg}(\text{OAc})_2$ the decompn. is so slow that at the end of 1 hr. the rotation is still rising because of the increasing concn. of sugar with high rotation, and only after 2 hrs. does the mutarotation begin to manifest itself by a decrease in the rotation. These results show plainly that the sugar split off is the α form, and would seem to confirm Wrede's conclusion that sinigrin and sinapin are α glucosides. The synthetic Na β glucosulfate should then give with AgNO_3 β glucose with a low initial rotation, but as a matter of fact there was obtained a glucose soln. with $[\alpha]_D^{20}$ changing from 68.5° after 45 min. to a const. value of 48.5° (that the final value is somewhat lower than the equl. value, 52.5°, for glucose is due to the fact that the colloidal glucosulfate Ag compd. is adsorbed on the $\text{BaCO}_3\text{-Ag}_2\text{S}$, if the decompn. is carried out in the absence of BaCO_3 the mutarotation is considerably accelerated but leads to const. values very close to 52.5°). In the desulfurization of β glucosulfate by AgNO_3 the glucose is therefore split off as α glucose, and as the reaction is entirely analogous to that with the mustard oil glucosides the latter are certainly β glucosides. This is confirmed by the formation of β , β -diglucosyl disulfide octaacetate, m. 140°, $[\alpha]_D^{20} -158.7^\circ$ in $(\text{CH}_3\text{CH}_2)_2$ from sinigrin heated on the H_2O bath with about 2.2 mols. of 0.1 *N* NaOH , cooled to room temp., neutralized, immediately treated with 0.1 *N* I as long as it is used up, evapd. *in vacuo* and treated with Ac_2O in cold $\text{C}_6\text{H}_6\text{N}$. C. A. R.

Digitalinum verum. H. KILLIAN. *Ber.* 63B, 2866-9 (1930).—For the hydrolysis of *Digitalinum verum* (I) K. had previously always used a dil. HCl contg. alc. and then treated the sugar soln., after removal of the digitaligenin, with Br to sep. the oxidation product of the digitalose as digitalonolactone (II) from the d gluconic acid formed at the same time, but the yield of II was very small (never more than 10%). It therefore seemed desirable to devise another cleavage method for such glucosides, especially since the use of HCl contg. alc. leads to the formation of considerable quantities of ethylsugars, which greatly interfere with the sepn. and identification of the individual sugars. K. had left, from his earlier work, several kg. of digitalis by-alkaloids which, he found still contained considerable quantities of I. This he extd. by a method similar to that used before but differing in the relative proportions of substance and solvents employed, he will publish the details of this method after he has tested it on a com. by product to be soon supplied to him. The 90 g. of crude I he obtained by his new method was purified by the method previously described but for its hydrolysis he used an acid contg. AcOH instead of alc. (35 cc. glacial AcOH , 55 cc. H_2O and 10 cc. concd. HCl), the solvent action of the AcOH on both the glucoside and the genus is

about the same as that of the alc, no ethylsugars are formed and, because of the diln. and the presence of HCl, the danger of acetylation is exceedingly small. Details of the hydrolysis and of the working up of the sugar soln for II are given. The yield of H is regularly around 40%. Attention is called to an error in Haworth's book, *The Constitution of Sugars*. C. A. R.

Strophanthin. XIX. The dehydrogenation of strophanthidin and gitoxigenin. WALTER A. JACOBS and ELMER E. FLECK. *Science* 73, 133-4 (1931), cf. C. A. 24, 5773.—The Se method of dehydrogenating is used and gives promising results pointing toward the structure of the hydrocarbon upon which these aglucones are built. Four fractions were obtained from strophanthidin and 2 from gitoxigenin. The hydrocarbon obtained from fraction II from gitoxigenin resembles the $C_{20}H_{32}$ hydrocarbon obtained from fraction 4 from strophanthidin. W. H. TIFFANY

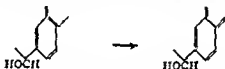
Unedoside, a new glucoside hydrolyzable by emulsin, from the fresh leaves and branches of *Arbutus unedo* L. M. BRIDEL and (MILIE) C. BOURDOUIL. *J. pharm. chim.* [8], 12, 241-53. *Bull. soc. chim. biol.* 12, 910-20 (1930) cf. Bourquelot and Fichtenholz, C. A. 4, 2945.—Unedoside (I), obtained with a yield of only 0.1 g. per kg. of fresh leaves or branches, is sol. in water (3% at ordinary temp.) and 95% alc. and insol. in Et_2O , crystallizes from Me_2CO contg. 7% water in long prisms, m. 225.5°, loses only 0.35% of its wt. in vacuo at 50°, $[\alpha]_D^{25} -112.43^\circ$, $[\alpha]_{589} -133.92^\circ$, reducing power, 0.853 that of dextrose, similar to that of verbenaldehyde (verbenalin, Bourdier, C. A. 2, 661, 1595), the latter differs, however, by a higher ro p. and lower $[\alpha]$. Upon hydrolysis with H_2SO_4 , I gives a black ppt. caused by the decomn. of the nonglucidic product *unedol*, the rotation becomes pos., but the reducing power diminishes by more than 20%. Following hydrolysis of I with emulsin, unedol does not form a black ppt. as with H_2SO_4 , but the soln. turns red brown, unedol ppts. and the reducing power of the soln. diminishes, hence I is the only glucoside whose hydrolysis is characterized by a loss of reducing power. Unedol has been crystd., it also possesses reducing properties. A bibliography of 8 references is added. S. WALDBOTT

Chlorophyll. XV. Phyllo- and pseudophylloerythrins. II. FISCHER, OTTO MOLDENHAUER and OSKAR SÖS. *Ann.* 485, 1-25 (1931); cf. C. A. 25, 961.—Phylloerythrin (I) gives with $MeOH-KOH$ a quant. yield of the *K* salt, $C_{40}H_{56}N_4O_8K$, sol. in H_2O but insol. in all org. solvents, 12% HCl gives I. Heating I with $m-C_6H_4(OH)_2$ gives a porphyrin spectroscopically identical with pyrrhoporphyrin. I in $AcOH$ -concd. HCl treated with H_2O_2 , gives a Cl deriv., probably a mixt. of $C_{40}H_{56}N_4O_8Cl$ and $C_{40}H_{54}N_4O_8Cl$. I and $H_2NCONHNH_2$, in CaH_2N yield the semicarbazone, small needles, $C_{40}H_{56}N_4O_8 \cdot N_2O_2$, for which the spectrum in CaH_2N is given, the max. is at 631.2, the *Me* ester, $C_{40}H_{56}N_4O_8$, forms deep red prismatic needles. $N_2H_4 \cdot H_2O$ and I give a *hydrazone*, prismatic needles with a red violet surface luster; the spectrum is shifted to the violet; after short exposure to light the spectrum of I is again seen. I also yields an *oxime*, $C_{40}H_{56}N_4O_8$, prismatic needles; the *Me* ester, m. 270° (decomn.), forms deep red needles. Reduction of I by heating with $EtONa$ and $N_2H_4 \cdot H_2O$ 6 hrs. at 165-70° gives *desoxyphylloerythrin* (II), $C_{40}H_{56}N_4O_8$, prisms or prismatic needles with blue-violet surface luster; HCl no. about 1.5, the spectrum is given in pyridine-ether and HCl ; *Me* ester, m. 264° (cor), *Fe* complex, $C_{40}H_{56}N_4O_8FeCl$, prismatic needles, *Cu* salt, tablets from $CHCl_3$, $MeOH$. Reduction of the semicarbazone of I with $EtONa$ at 168° gives a porphyrin spectroscopically identical with isopheoporphyrin a. Pseudophylloerythrin (III) with ice-cold oleum for 15 min. gives I; while 2 mols. I crystallizes with 1 mol. $CHCl_3$, III does not do so, even after heating with HCO_2H 24 hrs., showing that III is unchanged by this treatment. Oxidation of III with H_2O_2 gives rhodoporphyrin and a 2nd product, probably isochloroporphyrin c. III gives an *oxime* (*Me* ester) and a semicarbazone. The di-*Me* ester of diacetyldiuteroporphyrin yields a *di-oxime*, needles, and a semicarbazone, for which the spectra are given. Reduction of this ester gives a mixt. of hemo- and mesoporphyrins. II does not react with $EtONa$ and N_2H_4 at 170°. 3,4'-Dimethyl-4,3'-diethyl 5-bromo-5'-bromomethylpyrrromethene-HBr and 3,3',4',5' tetramethyl-4 propionic acid pyrrromethene-HBr, on fusion with $(CH_3CO_2H)_2$ at 210° give 1,3,5,6,7-pentamethyl 2,4-diethyl-8 propionic acid-porphin, whose *Me* ester, $C_{40}H_{56}N_4O_8$, m. 242° (cor), its spectrum is entirely different from that of II. The *Fe* and *Zn* salts of III are reported. C. J. WEST

The enzymic hydrolysis of gelatin in its relation to the formation of diacipiperazines. A. BLANCHETIÈRE, *Compt. rend.* 191, 1479-81 (1930); cf. C. A. 23, 1619, 2698, 4219.—A relation has been pointed out between the N of the diacipiperazines formed in fermentative hydrolysis of proteins and the N content of certain amino acids known to be resistant to hydrolysis. Gelatin differs much in chem. behavior from ovalbumin and gliadin. In peptic digestion, the ratio of diacipiperazine N in

gelatin as compared with that in ovalbumin is 0.84 while the corresponding ratio of amino acid to ovalbumin is 0.97. In comparative expts., the ratio of diacipiperazine N in gelatin to that in gliadin is 0.295 while the corresponding ratio of amino acid to gliadin is 0.471. This is interpreted to indicate that the diacipiperazine N is less in gelatin than in certain resistant anhydrides of amino acids, and also that gelatin has less N linked by cyclopeptidic structure than has ovalbumin or gliadin. N. M. N.

Ergosterol and some of its derivatives. Z. DE BOE. *Bull. acad. roy. m'ed.*, Belg. 10, 336-75 (1930), cf. Castille, C. A. 24, 6320.—A critical review, with 34 references, of the literature on the isomerization, reduction and oxidation derivs. of ergosterol A (I) and a study of certain isomerizations and reductions of L. "Iso-ergosterol" was prep'd. according to Reindel (cf. C. A. 21, 1461) and, confirming Heilbron (C. A. 24, 1388), shown to consist of a mixt. of *isoeergosterol* B (II) and *ergosterol* B (III), both of which possess a secondary alc. group and 3 double bonds and correspond to $C_{28}H_{46}O$. II, crystd. from MeOH-EtOAc (4:1) or EtOH, m. 135-6°, has a $-\text{CO}_2\text{H}$ and forms an *acetate* (IV), m. 112°. III m. 145-6° and turns on standing into II. Sapon. of its *acetate* (V), m. 139°, yields mainly II. The ultra-violet absorption spectra of II-V in $C_{12}H_{10}$ and EtOH-EtO are described. A mixt. of II and III was also obtained by the action on I of dry HBr at 0°, while by boiling I with H_2SO_4 in EtOH only II was formed. From a soln. of I in dry pyridine and $BzCl$ kept at 0°, *ergosterol benzoate*, m. 164-5° after recrystns. from MeOH and EtO, was isolated. Its absorption spectrum in EtOH and $C_{12}H_{10}$ is described. By heating I with $BzCl$ to 150-70°, *isoeergosterol benzoate*, m. 125-6° after recrystn. from MeOH-EtOAc, is formed. Sapon. with KOEt yields a mixt. of II and III. From the absorption spectra the conclusion is drawn that the isomerization of I into II and III consists of a shifting of double bonds such as



By refluxing I in AmOH in presence of Na and avoiding air oxidation, *dihydroergosterol* A, $C_{28}H_{48}O$ (VI), free of I and m. 208-9° after crystn. from dry MeOH-EtOAc or dry EtO, was obtained, while crystn. from ordinary EtOH gave the *dihydrate* of VI, m. 203° and retaining 2 H_2O even after 6 hrs. at 120° over H_2SO_4 . From the MeOH-EtOAc filtrates *monohydroergosterol* A, $C_{28}H_{46}O$ (VII), m. 173-4°, was isolated. Sapon. of the *acetate* of VI, m. 181°, gave VII. These relations are confirmed by a study of the absorption spectra. A similar reduction of III could not be realized, but when III was refluxed in AmOH in presence of Na, *α-isoeergosterol* B, m. 188-90°, possessing 3 double bonds and 1 secondary alc. group, could be isolated. Sapon. of its *acetate* m. 112° gave II. The corresponding absorption spectra are reported. It is shown that Heilbron's observations on the alleged isomerization of VI (C. A. 24, 623) by HCl gas are merely due to a m. p. depression caused by transformation of I, present in impure preps. of VI, into III; no isomerization being observed with pure VI. An attempt to reduce I by HI and red P at 230° led to a non-cryst. material free of alc. groups, evidently contg. only 1 double bond and possibly consisting of a mixt. of *ergostane*, $C_{28}H_{48}$, and *ergostene*, $C_{28}H_{46}$. The same mixt. seems to be formed by the same reduction of II.

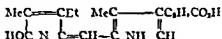
C. TOENIES

Soy-bean lecithin. II. Lecithins of the α-series. YOSHIXUNI YOKOYAMA AND BUNSUKE SATO. *Proc. Imp. Acad. (Japan)* 7, 12-4 (1931), cf. C. A. 25, 630.—The α-lecithins obtained as double salts with $CdCl_2$ were brominated and fractionated by means of petr. ether, H_2O and $CHCl_3$, 4 bromides were isolated, the first 3 being *only palmito-inoleo-α-lecithin*, *oleo-inoleo-α-lecithin*, *dioleo-α-lecithin* and *dilinoleo-α-lecithin*, m. 113-4°, these structures were established by hydrolysis and identification of the fatty acids.

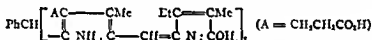
C. J. WEST

Neo-, xantho-neobilirubin acid and partial synthesis of mesobilirubin and mesobilirubinogen (urobilinogen). H. FISCHER AND RICHARD HESS. *Z. physiol. Chem.* 194, 193-228 (1931).—Bilirubin, the biol. degradation product of the blood pigment, contains 33 C, its cleavage product, bilrubic acid, obtained by energetic reduction, contains 17 C and represents, therefore, only 0.5 of the bilirubin mol. The other half has hitherto escaped identification. This unknown half of the bilirubin mol., *neobilirubin acid* (I), has now been isolated and identified. The method employed for the cleavage of bilirubin was the resorcinol fusion used by Schumm (C. A. 23, 153)

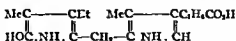
for breaking off the unsatd side chains of hemin. Bilirubin itself did not give satisfactory results under this treatment, but its hydrogenation product, mesobilirubin, yielded a yellow cryst product *neoxanthobilirubic acid* (II). The relationship between II and I is that of pigment and leuco compd. The synthesis of mesobilirubin was effected by condensing 2 mols of II with 1 CH_2O . This supplies the additional C atom of the bilrubic acid half of the mol, the other half being neobilrubic acid itself. *Neoxanthobilirubic acid* (II), m. 229° (Me ester m. 190°), was prepd by adding mesobilirubin in boiling resorcinol, pouring the molten mass into H_2O , extg the ppt with CHCl_3 and recrystg from MeOH. It contains a free CH since it yields a cryst *monoazo dye*, m. 193° (decompn), with PhN_3Cl . Bromination converts it into a *mono-Br deriv.*, darkens 171° but does not m below 350° in which the Br is substituted on a pyrrole nucleus. With $\text{CaH}_2(\text{CO})_2\text{O}$ at $180-90^\circ$ it forms an orange-yellow *phthalide*, m. 298° . On reduction with Na-Hg it yields the leuco compd I, which gives an intense Ebrlich reaction and on treatment with MeOH at 190° reverts to II. On oxidation it yields methylethylmaleic imide, m. 67° . It contains 1 CH_2 less than xanthobilirubin. The probable formula of II is



While bilrubic acid breaks down into cryptopyrrole and cryptopyrrolecarboxylic acid, a substance of the above structure should yield cryptopyrrole and heraopyrrolecarboxylic acid. Both of these products were actually obtained and identified. The pyrrole ring contains a reactive CH which condenses with aldehydes, 2 I. With BzH the product is *bis(neoxanthobilirubic acid)phenylmethane*, m. 257° , represented by the formula

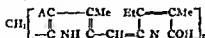


In the Gmelin reaction this product gives a sequence of colors similar to that of bilirubin. It is the first synthetic product to give this characteristic reaction. When reduced with Hf-AcOH it yields *benzylneobilrubic acid*, m. 187° . The other product should be I, but it has not yet been identified. Other aldehydes give similar condensation products with II; $\text{o-N}_2\text{NC}_6\text{H}_4\text{CHO} \rightarrow \text{bis(II)-o-nitrophenylmethane}$, m. 259° , $p\text{-Me}_2\text{NC}_6\text{H}_4\text{CHO} \rightarrow \text{bis(II)-p-dimethylaminophenylmethane}$, m. 239° , 4-methyl-3,5-diacethoxy 2 formylpyrrole \rightarrow condensation product, m. 239° , $\text{AcH} \rightarrow \text{bis(II)methylmethane}$, m. 267° , $\text{CH}_3\text{O} \rightarrow \text{bis(II)methane}$, m. 310° . This last product coincides with mesobilirubin in cryst. form, m. p., mixed m. p. and Gmelin reaction. It gives the same Me ester, m. 190° , which forms the same red HCl sol. Both mesobilirubin and the synthetic product are reduced by Na-Hg to the leuco deriv., *mesobilirubinogen*, both give an intense diazo reaction, and with $\text{Zn}(\text{OAc})_2$ yield *urobilin*, m. 192° , with its characteristic fluorescence. The crystallographic identity is complete. All of the above aldehyde condensation products are yellow and well crystd. Condensation products were also obtained with MeAc and with camphor. Reduction of II by Hf-AcOH or by Na-Hg gives the leuco compd *neobilrubic acid* (I), m. 185° . This gives a sharp m-p depression with bilrubic acid which m. 187° . Like II it contains a free CH and condenses with aldehydes. The condensation product with BzH , m. 248° , is identical with that obtained from mesobilirubin and BzH (Me ester m. 212°). The PhCH deriv. is reduced by Hf-AcOH to *benzylneobilrubic acid*, m. 189° . The latter cannot be brominated, but with Br-AcOH it undergoes oxidation to *benzylidene-neobilrubic acid*, m. 248° . These 2 substances represent pigment and leuco compd. Bilrubic acid, on the other hand, does not condense with BzH since the pyrrole rings are completely satd. The formula for I is

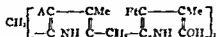


Reduction of mesobilirubin by Hf-AcOH gives both halves, I and bilrubic acid, in 50% yield. They may be sep'd by CHCl_3 in which the former is difficultly and the latter easily sol. The condensation products of I and II with aldehydes represent 2 distinct types, similar to those obtained with simple pyrroles, viz., the tripyrrolmethanes and the dipyrrolmethenes. MeAc condenses first to mesityl oxide and

this with II to bis(II)isobutylmethane, m 256°. Mesobilirubin is



and its leuco compd., urobilinogen,



The main difficulty in explaining the urobilin reaction is the fact that dehydrogenation does not give mesobilirubin. Air oxidation gives, instead, urobilin. Perhaps the formation of urobilin consists in dehydrogenation to a trimethin pigment in which numerous possibilities of isomerism exist. Such a structure would be extremely unstable and inclined to formation of resins. All this confirms the supposition that hemin consists of 4 pyrrole rings which are linked by 4 methin bridges. Hemin contains 34, bilirubin 33 C atoms. A connecting CH group thus steps out and is replaced by 2OH in the opened porphyrin ring, and the new bis(hydroxypyrr)omethin pigment, the hypothetical structure of urobilin but with $\text{C}_{31}\text{H}_{42}$ in place of $\text{C}_{33}\text{H}_{40}$ groups, would then occur as intermediate product in the formation of bile pigment. There still remains, however, the possibility of a more complete breakdown of the blood pigment and a resynthesis to bilirubin from the resulting fragments. A. W. DOX

The structure of quinoid compounds and of a molecular compound of the quinoxaline type (HERZEL, RÖMER) 2. The crystal structure of the diphenylpolyenes (HENGSTENBERG, KUTNY) 2. A determination of the molecular weight of methylthiurea by means of x rays (HENGSTENBERG, KUTNY) 2. A powder spectrometric study of urea (WYCKOFF) 2. Vapor pressures of some hydrocarbons (LINDER) 2. Studies in x ray diffraction. III. Aromatic hydrocarbons in the solid and liquid states (KAISER, NAMURTI) 3. The primary oxidation of bituminous coal (FUCHS, HORN) 21. Anodic oxidation of AcOH in H_2SO_4 solution (SCHREINER) 4. Polymerization and thermal decomposition of C_{12}H_8 (SCHLÄPFER, BRUNNER) 2. The dehydrating action of coal ash (TANAKA) 18. Frangulic acid (BAIDEL, CHARAUX) 11D. Composition and properties of certain red and yellow plant pigments (BILGER) 11D.

BODROUX, FERNAND. Cours de chimie organique à l'usage des candidats aux certificats d'études physiques, chimiques et naturelles. Paris: Vuibert. 242 pp.

WERTHEIM, EUGEN. Essentials of Organic and Biological Chemistry. Easton, Pa. The Chemical Pub. Co., Inc. 175 pp.

Condensing hydrocarbons. IMPERIAL CHEMICAL INDUSTRIES, LTD. Fr. 694,350, Apr. 23, 1930. Aromatic or unsat'd hydrocarbons are prep'd by passing hydrocarbons contg. at least 2 atoms of C per mol. through a heated space at a temp. of at least 1000° or 1100°, preferably 1200°, and at a volumetric speed of 50 or 100, preferably more. The treatment may be carried out in several steps with removal of aromatic hydrocarbons between the steps, the temp. and speed increasing from step to step. The walls of the reaction chamber may be made of agglomerated SiC or graphite or an alloy resistant to heat. If it is desired to favor the production of $\text{C}_{12}\text{H}_{10}$, the walls may be made of sillimanite.

Synthetic hydrocarbons. HENRY C. WADE. Fr. 694,324, Mar. 29, 1930. Liquid synthetic hydrocarbon products are prep'd by vaporizing crude petroleum or portions thereof having relatively high b. ps. or liquefied coal or like materials, and during this treatment leading the vapor through a spiral path with speed and without pressure, and through a zone contg. a metallic catalyst at a moment when the hydrocarbon reaches the atomic state and along with steam to produce nascent H and free C which combine. The vapors are condensed or fixed in the state of gas by a suitable heat treatment. An app. is described.

Hydrogenating aromatic hydrocarbons and their derivatives. I. G. FARREND. A.-G. Brit. 328,583, July 10, 1929. Comps. such as C_{12}H_8 , naphthalene or their hydroxy or amino derivs. are hydrogenated at temps. below 350° in the presence of mixed catalysts immune from poisoning from S and the like and which comprise oxides or hydroxides or sulfides of metals of group 6 of the periodic system in admixt. with

metals of group 8 or heavy metals of group 1 (as such or in the form of their compds.) deposited on carriers of large surface. Examples are given of the use of sulfides of Pt and of W deposited on active silica, oxides of Ni and Mo and CaCO_3 deposited on active charcoal, and sulfides of Co and Mo deposited on Florida earth.

Hydrocarbon derivatives. I. G. FARBERND. A-G. Fr. 695,095, May 5, 1930. Hydrocarbon derivs. contg. in the mol. more than one halomethyl group are prepd. by the reaction on polynuclear hydrocarbons, of an aq. soln. of CH_3O satd. with hydrocarbons. Examples are given of the prepn. of di(*o*-chloromethyl)naphthalene, m. 130–140°, chloromethyl derivs. of $\text{C}_{10}\text{H}_{14}$ and a product corresponding to $\text{C}_{10}\text{H}_{14}(\text{CH}_2\text{Cl})_2$, m. 117.5°.

Halogenation of paraffin hydrocarbon series. N. V. BATAAFSCHE PETROLEUM MAATSCHAPPIJ. Dutch 23,383, Jan. 15, 1931. For the halogenation of satd. aliphatic hydrocarbons an enameled reaction vessel is used, this prevents both corrosion and C liberation.

Halogenated isocyclic compounds. I. G. FARBERND. A-G. Brit. 335,948, April 3, 1929. Isocyclic compds. substituted in the nucleus by Cl or Br or both are prepd. by reaction on isocyclic compds. with HCl or HBr or both in oleum or chlorosulfonic acid or their mixts. in the presence of halogen carriers such as S, I, Sb, As, Bi, Hg or their compds. Temp. conditions may be chosen to avoid or favor side-reactions such as prepn. of halogenated sulfonic acids or halogenated quinones. Among the starting materials which may be treated are benzene, naphthalene, anthracene, pyrene, perylene and their nitro, amino and halogen derivs., benzanthrones, dibenzanthrones, *ms*-benzodanthrones, *ms* naphthodanthrones, *allo-ms*-naphthodanthrones, *ms*-anthradanthrones, anthanthrones, dibenzopyrenequinones or isomeric compds.

Converting olefins into higher olefins and aromatic hydrocarbons, etc. I. G. FARBERND. A-G. Brit. 336,234, May 4, 1929. Starting materials such as C_2H_4 , C_3H_4 , C_4H_4 , or gas mixts. contg. them such as casing head gases or gases produced in cracking mineral or tar oils are brought into contact with elementary Si or a Si-contg. material such as ferro-Si contg. a large proportion of Si, in the form of small pieces or bricks (suitably at 350–800° and under pressures which may be up to 1000 atm. or more). App. is preferably used which resists H_2S and which does not cause deposition of C, such as "V2A" steel or steels contg. Mo and W. Various details and examples are given.

Separating olefins. N. V. BATAAFSCHE PETROLEUM MAATSCHAPPIJ. Fr. 694,323, Apr. 23, 1930. The concn. of one constituent of a fluid mixt. is increased by effecting the fractionated distn. of the mixt. in the presence of a solvent having a selective attraction toward one of the constituents. Thus, to sep. a mixt. of olefin hydrocarbons and paraffins NH_3 is added. An example is given of the sepn. of butylene and butane.

Oxidation products from hydrocarbons, waxes, etc. I. G. FARBERND. A-G. Brit. 337,130, Sept. 12, 1929. An app. is described for producing oxidation products from solid or liquid hydrocarbons, waxes and the like by the action of oxidizing gases, in which cooling liquid is sprayed on the outer surface of the reaction vessel which is surrounded by a pressure-tight cooling jacket.

Oxygenated organic compounds from carbon oxides and hydrogen by catalytic reactions. H. DREYFUS. Brit. 337,014, July 24, 1929. Compds. such as EtOH, HOAc, acetaldehyde, MeOAc and higher homologs are made by use of catalysts comprising ferrites, ferrates, cobaltites or cobaltates such as those of alkali or alk. earth metals or of Al, preferably at temps. of 200–600° (suitably 250–350°) under 50–500 atm. or higher pressure. There may be used coke-oven gases, producer gas or water gas which may be preliminarily treated if desired to effect formation of synthetic products such as MeOH or EtOH in the presence of a catalyst such as ZnO or a basic chromate. The catalysts may be used in connection with Zn, Mg, Ca, Al, Cr, Mo, V, W or U or their oxides or other compds. Cu may be present, and the app. used may be of Cu or steel lined with Cu or of steel contg. Mn, W, Co or Ni.

Organic bases. I. G. FARBERND. A-M. (Otto Nicodemus and Walter Schmidt, inventors). Ger. 516,765, Sept. 27, 1928. Addn. to 479,351 (C. A. 23, 4709). Org. bases, generally of the pyridine series, are prepd. by passing C_2H_2 and the vapor of an aliphatic primary or secondary amine at 250–400° over a catalyst comprising an inorg. or org. salt capable of combining with NH_3 , or a metal oxide such as Fe_2O_3 , ZnO or Al_2O_3 . Examples are given.

Organic isocolloids. LASZLO AUER and LAJOS SUSZTEK. Fr. 694,339, Apr. 23, 1930. The phys. properties of org. isocolloids composed of unsatd. org. acids of high mol. wt. or contg. such acids, are modified by incorporating therein, by soln. or dispersion, soaps of alkali or alk. earth metals, Zn, Mg, etc., and besting to a relatively

low temp. The treated products may be vulcanized. Examples are given in which linseed oil, K or Na soap and linseed oil acids are heated to about 200°. An after treatment with S gives products which are useful in varnish and rubber manuf. Cf C A 25, 618

Organic isocolloids. LASZLO AVER. Fr. 694,340, Apr. 23, 1930. The soly. of modified or unmodified org. isocolloid substances is altered by dissolving or dispersing in such substances, agents favoring soln., *s. e.*, agents which are themselves sol. in the solvent in which it is desired to render the isocolloid sol. or more sol. Thus, castor oil is heated with NH_4I in *vacuo* for 5 hrs. at 200° to give a dark colored soft paste which is sol. in acetone. Other examples are given.

Organoarsenic compounds. I. G. FARBENFABR. A.-G. Brit. 337,299, Feb. 11, 1929. *N*-Alkyl-, aralkyl-, or -alkylene-3,4-benzimidazolone-*o*-arsonic acids are prep'd by treating *o*-aminoalkylamino-, *o*-aminoaralkylamino- or *o*-aminoalkyleneamino-benzene-*o*-arsonic acids with a chloroformic acid ester and decompg. the product with an acid (preferably with heating). Details are given of the production of 1-methyl-, 1-propyl- and 1-allyl-2,3-dihydro-2-ketobenzimidazole-5-*o*-arsonic acids by interaction with 1-*l*-chloroformate of 4-methylamino-, 4-propylamino- and 4-allylamino-3-amino-benzene-*o*-arsonic acids, resp.

Separating organic liquids. GENERAL TECHNICAL CO., LTD. Fr. 695,005, Aug. 2, 1929. The constituents of mixts. of org. liquids of different b. ps. are sep'd by passing the mixts. through a tube or tubes, preferably filled with contact material, a counter-current of vapors, of a non-miscible liquid being circulated, the temp. of treatment being slightly higher than the b. p. of the non-miscible liquid. An example is given of the sep'n. of a mixt. of hydrocarbons using a counter-current of steam.

Catalysts for reducing or dehydrogenating organic compounds. I. G. FARBENFABR. A.-G. (Hellmut Langbeinrich, inventor). Ger. 516,251, May 7, 1926. Oxygenated compds. of Mo or W are treated with NH_3 (or gases contg. NH_3) at 450-600° and at atm. or raised pressure. The compds. obtained from molybdates or tungstates of Cu, Ag, Fe and Ni are particularly suitable. Examples are given.

Alcohols. I. G. FARBENFABR. A.-G. Fr. 694,424, Apr. 25, 1930. Alcs. are prep'd by treating sugars with H₂ in the presence of catalysts other than Pt and in the presence of OH ions. Examples are given of the prep'n. of sorbitol and mannitol from glucose in the presence of activated Ni and CaO, etc.

Alcoholates. CHRISTIAAN VAN LOON. Ger. 513,577, July 22, 1926. Alkali metal alcoholates of polyhydric alcs. are prep'd by evap'g aq. solns. of the alc. and an alkali hydroxide at high temp. and reduced pressure. The reaction is preferably carried out in a current of indifferent gas. Thus, a soln. contg. 1 mol glycol and 1 mol conc'd NaOH is evap'd in a water bath in *vacuo*. The product contains 20-25% Na. Other examples are given.

Phenolic aldehydes. SCHWABE & CO. A.-G. Ger. 513,578, Oct. 18, 1925. The ozonides of the corresponding propenyl compds. are reduced by hyposulfite. Thus, isoeugenol in CHCl_3 is sat'd. with O_3 , and the resulting ozonide treated with $\text{Na}_2\text{S}_2\text{O}_4$ to give an 84% yield of vanillin. Other examples are given.

Esters. I. G. FARBENFABR. A.-G. (Erwin Schwabe, inventor). Ger. 516,135, Aug. 6, 1925. In the manuf. of esters from acids that are relatively non-volatile and alcs. that are insol. or difficultly sol. in water, a reaction mixt. is used contg. only a small excess, if any, of acid or alc. and also contg. a water insol. solvent for the alc., preferably a solvent b. near 100°. The distillate obtained from the mixt. sep's into 2 layers, the aq. layer being removed and the layer of solvent and unreacted alc. being returned to the reaction vessel. The prep'n. of isobutyl phthalate with the use of toluene as the solvent is described in an example. Cf C A. 24, 2141.

Esters of carbonic acid. I. G. FARBENFABR. A.-G. (Gerhard Steimmg and Max Wittwer, inventors). Ger. 516,281, Nov. 15, 1928. Carbonic esters of glycols are prep'd by treating α , β -chlorohydrins with alkali carbonates or bicarbonates. Thus,

glycol carbonate, $\text{CH}_2\text{OC(O)OCH}_2$, may be prep'd by warming ethylenechlorohydrin 1 mol with NaHCO_3 1 mol for a few hrs. while stirring. Other examples are given also.

Esters such as fatty acid glycerides. I. G. FARBENFABR. A.-G. Brit. 336,276, July 4, 1929. In the production of esters of polyhydric alcs. such as glycerol with acids such as those of linseed or olive oil, the component of lower b. p. is led in vapor or mist form (and suitably with a carrier gas such as H_2 for simultaneous hydrogenation, N or H_2O vapor) at reduced pressure into the liquid higher-boiling component at a temp. above the b. p. of the lower-boiling component, and catalysts may be used such as Mg oleate, H_3PO_4 or ZnCl_2 . Cf C A. 25, 115.

Vinyl ethers. I. G. FARBEINEND. A.-G. (Otto Ernst and Walter Berndt, inventors). Ger 513,679, May 24, 1927. Vinyl ethers of aromatic hydroxy compds. of a phenolic character, without substituents with OH groups linked to aliphatic residues, are obtained by treating the aromatic compds. with vinyl halides, especially the chloride, in the presence of bases and, optionally, catalysts. Thus, PhOH, NaOH, water and vinyl chloride are fused together at 180–200° to give an 80% yield of vinyl phenyl ether of sp gr 0.977 (20°) and b 155.6°. Other examples are given. Cf. C. A. 25, 302.

Ethers of α -naphthylmethyl alcohol. I. G. FARBEINEND. A.-G. (Gustav Reddellen and Hans Lange, inventors). Ger 516,280, Apr. 3, 1929. α -(Chloromethyl)naphthalene is treated with aliphatic or aromatic alcs. in the presence of an acid binding agent. Examples are given of the prepn. of methyl α -naphthylmethyl ether, b₁₁ 134°, ethyl α -naphthylmethyl ether, b₁₁ 144.5°, butyl α -naphthylmethyl ether, b₁₁ 167°, and benzyl α -naphthylmethyl ether, b₁₁ 220°. Cf. Ger 508,890 (C. A. 25, 716).

Hydrogenating phenols. I. G. FARBEINEND. A.-G. Brit. 336,616, April 17, 1929. Hydroaromatic hydrocarbons are prepd. from phenols by hydrogenation at temps. of 250–400° under pressures above 50 atm. with catalysts comprising oxides or sulfides of metals of the second to seventh periodic groups (preferably those of the sixth group) in admixt. with Cu, Ag, Au or metals of the eighth group or their oxides or sulfides. Numerous details and examples are given.

Dioxanes. I. G. FARBEINEND. A.-G. Ger 516,844, Dec. 28, 1927. Addn. to 500,223 (C. A. 24, 4307). Single or mixed homologs of dioxane or mixts. of dioxane therewith are prepd. as described in Ger. 500,223, using as initial materials the mixts. of glycols obtainable from cracking gases, the by products obtained in etherifying such mixts., or the single higher glycols isolated from the mixts. Examples are given. Cf. C. A. 25, 524.

Indanones. I. G. FARBEINEND. A.-G. (Fritz Mayer, inventor). Ger. 515,110, June 11, 1927. Addn. to 513,204 (C. A. 25, 1260). Indanones are prepd. from xylenes, halotoluenes or dihalobenzenes by condensing these with a halide of β -chloropropionic acid, and treating the resulting ketone as described in Ger 513,204. Thus, 7-methyl-6-chloro-3-indanone, b₁₁ 188°, is obtained from *o*-chlorotoluene. Other examples are given also.

Naphthenates of alkaline earth and heavy metals. I. G. FARBEINEND. A.-G. Brit. 335,863, Feb. 22, 1929. In the prepn. of naphthenates such as those of Co, Pb, Mn, Zn, Ca, Pb-Mn or Co-Zn, naphthene acid is saponified with KOH and the resulting soln. is treated with a suitable alk. earth or heavy metal salt (all operations being carried out at boiling temp.) and the ppt. thus formed may then be dried at 130° in the same vessel.

Dialkyl sulfates. ÉTABLISSEMENTS LAMBIOTTE FRÈRES. Brit. 336,681, June 25, 1929. For making dialkyl sulfates, the corresponding ether (derived from the alc. such as MeOH or PrOH and H₂SO₄) is caused to react with oleum, preferably in theoretical proportions and at temps. below 50° (suitably 30°). Various details of procedure are given.

Sulfonamides. IMPERIAL CHEMICAL INDUSTRIES, LTD. N. BENNETT, H. DODD and W. C. SPRENT. Brit. 336,512, Jan. 17, 1930. Monoalkylxylenesulfonamides are made by direct alkylation of xylenesulfonamide with MeCl or EtCl under pressure in the presence of caustic alkali. Various details of procedure are given.

Alkali salts of aromatic sulfohaloamides. CHEM. FAB. VON HEYDEN A.-G. (Curt Philipp, inventor). Ger 515,465, June 20, 1928. Addn. to 514,094 (C. A. 25, 1261). The reagents necessary for the method of Ger 514,094 are mixed in the dry state, and made into tablets if desired. The reaction is effected by dissolving the mixt. in water. One mol. of the alkali may be dispensed with by using the sulfonamide in the form of its monosodium salt.

Aluminum gluconates. CHEMISCHE FABRIK VORM. SANDOZ. Brit. 336,922, Feb. 15, 1929. Sol. Al compds. forming stable solns. are obtained by reaction of Al compds. such as the sulfate with gluconic acid or its salts such as alk. earth metal gluconates (preferably in the presence of alk. earth metal hydroxides). The product may be evapd. to dryness or pptd. with alc. and the acid salts obtained may be neutralized with NH₃ or other alkali, or a neutral salt may be obtained directly by reaction of an alkali aluminate soln. with gluconic acid. Cf. C. A. 24, 5437.

Lead alkyl compounds. STANDARD OIL CO. OF ILL. Ger. 505,688, Aug. 27, 1926. A Pb-Na alloy is heated with the desired dialkyl sulfate with free halogen or a metal halide as a catalyzer. Thus a Pb-Na alloy contg. 10% Na is heated with (Et)₂SO₄ in toluene with KI in water as the catalyst, to give Pb(Et)₂, on distn.

Metal carbamates. I. G. PARBENIND. A-G. Brit. 336,740, Sept. 19, 1929. Reaction is effected between NH_3 , carbamate and finely divided (preferably dry) metal oxides such as those of Zn, Mg or alk. earth metals, in the absence of solvents, at temps. not usually substantially above 100° (suitably about $60-80^\circ$ in most cases), and if desired in the presence of a stream of an inert gas such as N_2 , air or N. The alk. earth metal carbamates thus produced may be converted into cyanamides by heating to about 700° and treating with a stream of NH_3 .

Pyrazolone derivatives. I. G. PARBENIND. A-G (Hans Grotowsky, inventor) Ger. 515,782, Apr. 5, 1927. Addn. to 514,421 (C. A. 23, 1684). The method of Ger. 514,421 is modified by using, instead of a prepd. benzylidene imine or a stoichiometric mixt. of the corresponding aldehyde and amine, a mixt. contg. a substantially lower proportion of the amine. 1 example is given.

Triazine derivatives. I. G. PARBENIND. A-G. Fr. 604,756, Apr. 30, 1930. See Brit. 334,887 (C. A. 25, 1266).

Hexahydrodiphenylamine compounds. I. G. PARBENIND. A-G. Brit. 337,103, Aug. 21, 1929. Two- or three-nuclear condensation products, obtained as described in Brit. 313,421 (C. A. 24, 1123) from cyclohexanone or its alkyl derivs. and aromatic amines or their products of acylation, are subjected to catalytic hydrogenation under pressure (with a base metal catalyst such as Ni and preferably in the presence of a diluent such as decahydronaphthalene). Details are given of the production of 4-amino-hexahydrodiphenyl from 4-aminotetrahydrodiphenyl, 4-amino-3-methylhexahydrodiphenyl from 4-amino-3-methyltetrahydrodiphenyl and of similar reactions, and catalysts are described contg. Ni, Co, Ce, Bi and Cu in various mixts. which may be carried on pumice or silica gel.

Condensation products. I. G. PARBENIND. A-G. (Arthur Löttinghaus, Heinrich Neresheimer and Wilhelm Schneider, inventors) Ger. 513,530, Feb. 23, 1929. Addn. to 487,870 (C. A. 24, 1807). Condensation products are obtained by converting benzanthrone, or derivs. with a free or halogenated Br. 1 position, or derivs. with S- or O-contg. substituents into the corresponding ketoaldehydes and then condensing these with other benzanthrone ketoaldehydes, benzanthrone or derivs. Thus, benzanthrone in PhNO_2 is treated with PCl_5 and concd. H_2SO_4 . The benzanthrone ketoaldehyde obtained is then condensed with benzanthrone in $\text{C}_2\text{H}_5\text{Cl}$. Other examples are given. Cf. C. A. 25, 1263.

Condensation products. HEYKEL & CRE. G. M. B. H. Ger. 513,797, Oct. 14, 1927. In condensing polyhydric alcs. and their ketonic derivs. with aliphatic, aromatic, mixed or hydroaromatic ketones, by heating to a high temp., the reaction is carried out without addn. of solvents, and in the presence of a small quantity of aq. or gaseous mineral acid or acid salt. Examples describe the condensation of glycerol with cyclohexanone and methylcyclohexanone. Cf. C. A. 24, 5044.

Halogenated fatty acids. I. G. PARBENIND. A-G. Brit. 336,623, July 15, 1929. Saturated or unsaturated fatty acids contg. more than 8 C atoms per mol. and which may contain OH groups are treated with halogens or with a mixt. of a chlorate and HCl to introduce at least 3 halogen atoms per mol. one at least of which replaces a H atom by substitution. Examples are given of the production of trichlorinated, tetrachlorinated and hexachlorinated ricinoleic acid, trichlorinated oleic acid, pentachlorodihydroxy stearic acid, heptachlorinated ricinoleic acid, tetrachlorinated oleic acid and hexachlorinated stearic acid, etc. The products are oils, usually having adhesive properties, or plastic or resinous materials. Cf. C. A. 24, 3517.

Aromatic dicarboxylic acids. RÖTGERSWERKE. A-G (Johannes Selisch and Ernst Guttman, inventors) Ger. 516,282, May 5, 1929. The oxidation of aromatic α -diketones to dicarboxylic acids, e. g., the oxidation of phenanthraquinone to diphenic acid, is effected in aq. alk. suspension with H_2O_2 , or in aq. suspension with Na_2O_2 or like metal peroxide. An addn. of MeOH or EtOH accelerates the reaction. Examples are given.

Sulfonic acids of the fatty series. H. TH. BÖHMER. A-G. Fr. 694,692, Apr. 30, 1930. Sulfonic compds. of higher fatty acids and their alkali salts are prepd. by treating the corresponding α -bromo fatty acids with an alkali sulfite. Thus brominated lauric acid is dissolved in aq. NH_3 , and the soln. is boiled under reflux with a concd. aq. soln. of $(\text{NH}_4)_2\text{SO}_3$, giving the NH_4 salt of the α -sulfolauric acid. The products are wetting, purifying, emulsifying and foaming agents.

Sulfonic acids, naphthenic acids, etc. N.-V. DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ. Fr. 694,578, Apr. 29, 1930. Purified sulfonic, naphthenic and like acids are prepd. by dissolving the initial material in water and treating with an amt. of tar oil, etc., sufficient to permit the substances to be purified, to sep. with the tar oil, after

which the aq. soln. is decanted and the tar oil is removed from the purified products by distn. Cf C A 24, 629

Aliphatic acids and esters such as acetic acid and methyl acetate. BRITISH CELLULOSE, LTD., and W. BAKER. Brit. 337,053, July 27, 1929. MeOH or other liquid aliphatic alc. is mixed with one or more acid catalysts such as H_3PO_4 and the mixt. is subjected to the action of CO at 200–450° and at pressures up to 300 atm. Esters in the reaction products may be saponified with H_2O in the presence of H_2SO_4 (or, preferably, H_3PO_4) at 150–300° to produce an acid and an ether. App. and various details of procedure are described.

Concentrating acetic acid. MARTIN MUGGAN and JOSEF WIMMER (to Consortium Elektrochemische Industrie, G. m. b. H.). U. S. 1,792,113, Feb. 10. See Fr. 671,283 (C A. 24, 2143)

Concentrating aqueous solutions of formic and acetic acids. JAMES NELSON, LTD. Ger. 513,574, July 13, 1927. See Brit. 281,827 (C A. 22, 3669)

Acetic and formic acid. IMPERIAL CHEMICAL INDUSTRIES, LTD. Fr. 694,016, Apr. 16, 1930. Aq. solns. of AcOH or HCOOH are concd. by passing the vapors thereof over active C at a temp. of 120–130°, whereby the acid is absorbed in preference. The absorbed acid is sep'd. from the C by raising the temp. and reducing the pressure.

Carboxylic acid chlorides. FRANZ HENLE (to I. G. Farbenind. A.-G.). U. S. 1,792,163, Feb. 10. In prep'g. carboxylic acid chlorides such as acetyl chloride, chloroacetyl chloride, propionyl chloride or butyryl chloride by reaction of the corresponding free acids with sulfochloride compds. such as Na chlorosulfonate or *p*-toluenesulfochloride, the reaction is rendered more efficient and almost quant. yields are obtained by adding to the reaction mixt. a "solid extender" such as NaCl, and other substances may be added such as a neutral sulfate, a pyrosulfate, kaolin or sand. Several examples with details of procedure are given.

Acetic anhydride and other aliphatic anhydrides. H. DREYFUS. Brit. 336,960, July 22, 1929. For producing the anhydride, the vapor of the corresponding acid is contacted with an oxide or acid of As or Sb or a salt of such an acid such as arsenates or antimonates of Na or K or Mg arsenate, which may be placed in tubes on a carrier. Various details and modifications of procedure are described.

Ethylene from acetylene. I. G. FARBENIND. A.-G. Brit. 336,999, May 24, 1929. Te or a compact form of Ni_3 is used as catalyst in forming C_2H_4 by hydrogenation of C_2H_2 . The Te may be used either in compact form or distributed on a carrier. Cf C A. 24, 4523.

Absorbing ethylene in sulfuric acid. N.-V. DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ. Brit. 336,603, June 4, 1929. Various compds. of Os, Ir, Pt, Ru, Rh, Pd, Cu, Fe, Co or Ni are used in the H_2SO_4 as catalysts in absorbing C_2H_4 or treating crude gases contg. C_2H_4 such as natural gas or cracking gases, to obtain a soln. which may be hydrolyzed, distd. or otherwise treated for the production of esters, alcs., ethers, etc. Various compds. of the catalytic metals may be added to the soln. and then converted into sol. complex compds. by passing CO or nitric oxide through the soln. Solvents such as EtOH, HOAc, $PhNO_2$ or ether may be added to the H_2SO_4 and froth-forming or emulsifying agents also may be added. Brit. 336,604 relates to a process of generally similar character for the absorption of olefins contg. more than 2 C atoms such as propylene or diallyl or crude gases contg. these or similar compds. Various details and modifications of procedure and examples are given.

Absorbing ethylene in sulfuric acid. N.-V. DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ. Brit. 336,633, June 4, 1929. The absorption is catalyzed by the presence of one or more of the metals Pt, Pd, Os, Ir, Ru or Rh or their compds., in finely divided form and preferably on a carrier such as carbon black, silica gel or decolorizing clay (suitably in the form of a colloidal ppt.).

Dichloroethylene. CONSORTIUM FÜR ELEKTROCHEMISCHE INDUSTRIE G. M. B. H. Fr. 694,054, Apr. 16, 1930. Dichloroethylene is prep'd. by passing trichloroethane over heated catalysts composed of salts of heavy metals such as Fe or Cu or alk. earth compds. such as $BaCl_2$.

Oxidizing propylene. IMPERIAL CHEMICAL INDUSTRIES, LTD. Fr. 694,726, Apr. 29, 1930. Isopropyl alc. and isopropyl acetate are obtained by absorbing propylene in a mixt. of concd. H_2SO_4 and glacial AcOH and hydrolyzing the sulfonic esters formed.

Methanol. H. DREYFUS. Brit. 335,962, June 5, 1929. In synthesizing MeOH from CO and H_2 , the reaction gases are passed first over one or more masses of ZnO with or without Cr oxide and then over one or more MeOH catalysts which are sensitive to S poisoning such as catalysts contg. Cu or Mn or their compds. Various details of procedure are described and water gas may be used as an initial material.

Methanol syntheses. ALWIN MITTASCH, MATTHIAS PIER and CARL MÜLLER (to I G Farbenind A-G) U. S. 1,791,568, Feb 10 In synthesis of MeOH from H and C oxides in the presence of a catalyst such as potash lime or potash-alumina more than 2 vols of H for each vol of CO and more than 3 vols of H for each vol of CO₂ are employed Several examples are given Cf C. A. 25, 715

Butanol and higher alcohols. N.-V. DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ. Brit 330,811, Feb 17, 1930 For producing BuOH and higher alcs, EtOH vapor under pressure (preferably 130-300 atm) is heated to not over 400° (preferably not much over 325°) in the presence of a catalytic mass contg MgO together with a smaller quantity of Cu oxide

Tetrahydrofurfuryl alcohol. E. I. DU PONT DE NEMOURS & Co. and IMPERIAL CHEMICAL INDUSTRIES, LTD Brit. 337,290, Jan 25, 1930 Furfural is heated under pressure in the presence of H, a Ni catalyst and water (suitably at 50-150° and under 100 lbs per sq in pressure) The catalyst is preferably prepd by pptg a soln. of Ni nitrate with (NH₄)₂Cr₂O₇ and NH₃, and partially reducing the ppt. with H at 600-650°.

Formaldehyde. BAKELITE CORP. Brit 336,282, July 10, 1929. In the catalytic oxidation of MeOH, formation of by-products is restricted by adding a basic agent such as NH₃, to the reaction mixt at any stage of the process prior to the distn of the crude product obtained by absorbing or condensing the mixt leaving the catalyzers, so that the distillate is rendered substantially neutral Brit 336,283 relates to the inhibiting of by product formation in similar processes by rapidly cooling the hot gaseous product leaving the catalyzers by bringing it into contact with an aq. liquid (preferably so proportioned that the temp. of the liquid does not exceed 50°) and the concn. of the formaldehyde is kept substantially below 30% Various details of procedure are described Cf C. A. 24, 866

Acetaldehyde. GUTENHOFFNUNGSGESELLSCHAFT OBERHAUSEN A.-G. Fr. 604,382, Apr. 24, 1930 AcH is prepd by passing ClH and CO₂ through an elec. field produced by elec. currents of high frequency and tension Contact substances may be placed in the field

Urea. A. B. LAMB. Brit. 335,913, June 27, 1929 In a cyclic process, liquid NH₃ and liquid CO₂ are continuously and separately introduced into a heated autoclave maintained under pressure, and the melt contg urea is continuously withdrawn from the autoclave, unconverted NH₃ and CO₂ are sep'd from the melt by distn and are sep'd from each other, and after being dried, compressed and liquefied they are reintroduced in liquid form into the autoclave Various details of app and procedure are described and producer gas (with conversion of its CO into CO₂) may be used as a starting material (the NH₃ being produced by catalytic synthesis from N and H). Cf C. A. 25, 305.

Thiourea. I G FARBENIND A.-G. Brit. 336, 111, Oct. 19, 1929. Ca cyanamide is treated with H₂S in the absence of liquids or in the presence of small quantities of liquids sufficient to form a paste (such as water or MeOH, EtOH, aniline, alkyl amines, EtOAc or other esters, pyridine or hydrocarbons such as C₆H₆ or their halogen derivs) and in some cases alk. substances such as NH₃ and gaseous amines are added. The materials may be subjected to high pressures and to stirring and the temp. usually should not exceed 100°. The thiourea is extd from the reaction mixt. by a selective solvent such as acetone, pyridine or its homologs, mixts of alcs and ethers or of alc and C₆H₆, or of alcs and CHCl₃, CCl₄ or C₂H₅, or in some case by aq. mixts. such as aq. alcs Various details of procedure are described Cf C. A. 24, 4524

5-Halo-2-amino-1-alkoxy- and -1-aralkyloxybenzenes and intermediates. WILHELM FITZKY (to General Aniline Works). U. S. 1,792,156, Feb 10 By the gradual addn of halogenating agents such as SO₂Cl₂ or Br₂ sufficient for the introduction of 2 halogen atoms, to suspensions (suitably in C₆H₅Cl or PhNO₂) of *o,o'*-dialkoxy- and *o,o'*-diaralkyloxydiphenylureas (obtainable by causing the corresponding amino compds. to react with phosgene in the presence of an acid binding agent) there are formed di-halogen compds of the general formula 4,2 X(RO)C₆H₃NHCONHC₆H₃(OR)X-2,4 wherein R stands for alkyl or aralkyl and X stands for halogen These compds., when split up by means of a caustic alkali or NH₄OH are converted into 5-halo-2-amino-1-alkoxy- and -1-aralkyloxybenzenes such as 5-chloro-2-amino-1-methoxybenzene, m. about 52°, 5-bromo-2-amino-1-methoxybenzene, m. 50-1°, 5-chloro-2-amino-1-ethoxybenzene, m. 32°, 5-chloro-2-amino-1-butoxybenzene, b.p. 180°, and 5-chloro-2-amino-1-benzyloxybenzene, m. 46-7°. Details of procedure for making all these compds are given.

1-Phenyl-2-methylamino-1-propanol (ephedrine). KNOLL A.-G. CHEMISCHER FABRIKEN and W. KLAVERN Brit. 336,412, July 30, 1929 1-Phenyl-1,2 propane-

dione (acetylbenzoyl) is treated with a reducing agent (suitably activated Al in the presence of water) in the presence of methylinimine

Flavanthrone, etc. W. SMITH, L. J. HOOLEY, J. THOMAS and SCOTTISH DYERS, LTD. Brit 336,943, April 16, 1929. 1,1'-Dianthraquinonyls contg Cl, Br or sulfonic groups in the 2,2'-positions are condensed with substances contg reactive H atoms such as NH_3 , monomethylaniline, aniline, aminonanthraquinones, phenols, mercaptans or Na phenate or similar compds. When NH_3 is used condensation is followed by ring closure of the 2,2'-diamino 1,1'-dianthraquinonyl to form flavanthrone. Various examples are given.

4,10-Diaminoperylene. F. BENSA. Brit 336,141, Dec 22, 1928. See Fr. 686,213 (C. A. 25, 525)

Isatin and its derivatives. GEORG KÄNZLPER, ARTHUR WOLFRAM and EMIL HAUSDÖRFFER (to General Aniline Works) U. S. 1,792,170 Feb 10. Cyanoformylides of the general formula RNHCOCN in which R stands for a substituted or unsubstituted aromatic residue having at least one free α position to the imino group are transformed into isatins by ring closure with a condensing metal chloride such as AlCl_3 or ZnCl_2 . Examples are given of the treatment of cyanoformanilide, cyanoform(2-methyl-3-chloroanilide) (yielding 6-chloro-7-methylisatin), cyanoform(3,5-dichloroanilide) (yielding the 4,6-dichloroisatin), *N*-cyanoformyl(β -naphthylamine) (yielding a 2,1-naphthylisatin, m 219°), cyanoform(3,5-dimethylanilide) (yielding 4,6-dimethylisatin), and *N*-cyanoformyl(1-chloro-2-naphthylamine) (yielding 1-chloro-2,3-naphthylisatin, m 262°).

Benzyl chloride. IMPERIAL CHEMICAL INDUSTRIES, LTD. Fr 694,429, Apr. 25, 1930. $\text{C}_6\text{H}_5\text{CH}_2\text{Cl}$ is obtained by treating the crude liquids from the manifold of benzyl cellulose with HCl . The liquids are given a preliminary treatment with an alk substance such as Na_2CO_3 to remove Fe or other metals. After removal of $\text{C}_6\text{H}_5\text{CH}_2\text{Cl}$ the residue is treated with Cl to give a mixt of $\text{C}_6\text{H}_5\text{CH}_2\text{Cl}$ and HCl .

Vinyl chloride. I. G. FARBENIND. A. G. Fr 694,575, Apr 26, 1930. Vinyl chloride is prepd by treating ethylene chloride with KOH dissolved in MeOH.

11—BIOLOGICAL CHEMISTRY

PAUL H. BOWE

A—GENERAL

FRANK P. UNDERHILL

Studies in the physical chemistry of muscle globulin. II. Some physico-chemical properties of muscle globulin [myosin]. JOHN T. EDSELL. *J. Biol. Chem.* 89, 289-313 (1930); cf. C. A. 22, 2174.—The muscle globulin is prepd by extg fresh, finely ground muscle with an excess of a 1.2 M KCl soln buffered with KH_2PO_4 and $\text{K}_2\text{H}_2\text{PO}_4$. The soln should have an ionic strength of 1.2 to 1.5 and pH 5.1 to 5.5. The protein is collected and purified by changing the ionic strength of the ext. several times, thus causing either a redissolving or repptn. When pptd, muscle globulin settles out as micellar particles which come to a thick jelly like mass on centrifugation. The pptd. protein has its isoelec. point at pH 5.1 to 5.5 and between pH 6.2 and 6.6 is the zone of minimal acid- and base-binding capacity. Myosin is almost completely insol. in the absence of salts or even in the presence of salts at low ionic strengths. As the phosphate soln approaches pH 7.4 there is a sudden rise in soly. between ionic strength 0.25 and 0.30. Myosin can be sep'd from serum globulin by pptn. in dil. salt solns. Myosin requires greater concns. of neutral salts to ppt. it than pptd. fibrinogen but somewhat less than is needed by eglobulin. At pH 7 and 0° it is completely pptd. by 4.4 to 4.6 M NaCl. Increasing the temp. decreases the soly. of myosin in concd. salt solns. In the absence of salts myosin is insol. from pH 4.5 to 8 and at all salt concns. It is insol. from pH 5 to 6. In moderate salt concns. it is sol. if the pH is alk. to 6.6 and as the acidity becomes greater than pH 5.5 the protein dissolves without the presence of a neutral salt. Alk. to pH 10, it is also sol. in the absence of salt. The viscosity of myosin is of a higher order of magnitude than that of the blood proteins of myogen. Even concd. ppts. of proteins contain 93 to 99% H_2O and this H_2O can not be removed unless the protein is denatured in the process. In the nearly complete absence of salt, and in the presence of small quantities of alkali, gels of the protein of a peculiar character are formed. Finally myosin is compared to the proteins studied by other investigators.

J. R. ADAMS

Studies in the physical chemistry of muscle globulin. III. The anisotropy of myosin and the angle of isochline. ALEXANDER L. VON MURALT AND JOHN T. EDMALL. *J Biol Chem* 89, 315-50 (1930); cf C. A. 22, 2174 and preceding abstract.—Highly anisotropic protein in myosin produces double refraction of flow. This is brought about through the orientation and elastic deformation of these particles. The app for the detn of the angle of isochline and the double refraction is described. The age of the myosin has no effect on the angle of isochline. For intermediate and high concns. at low temp it is between 77.5° and 78.5° and approaches 45° as a lower limit with increasing diln. For const temp and varying concn or const. concn and varying temp characteristic curves are obtained for the relation between the angle of isochline and angular velocity. Myosin soln possesses rigidity which is not affected by age. The gel formed from myosin is a thixotropic gel and gives two different crosses of isochline, one with an angle of 45° and the other 61° to 63° . The measurements of angle of isochline have been interpreted as an indication that the myosin particles are of uniform shape and size. IV. The anisotropy of myosin and double refraction of flow. *Ibid* 351-53.—A description is given of the app used to prove that myosin solns. contain anisotropic particles and that anisotropy is connected with the chem. structure of the protein. Double refraction and angular velocity are found to be related by characteristic curves, which vary with protein concn, are reproducible for any given prepn over a period of time and there appears to be no aging effect. The general form of these curves is the same for all preps examd. Denaturing agents produce rapid and complete destruction of the double refraction of flow. Double refraction is ascribed primarily to the orientation of anisotropic protein particles, due to the shearing stresses which arise during flow. Three tentative hypotheses are suggested to explain the observed facts. The Gans effect also indicates that the protein particles are anisotropic. The evidence also indicates that the anisotropic protein particles are of uniform size and shape. The properties of myosin solns. suggest that they may contain rod shaped particles. The double refraction of myosin would indicate its probable location in the anisotropic disk of the cross striated muscle fiber, and its general physicochem. properties suggest that it may play a part in the functional activity of the muscle. J. R. A.

Light and catalase. II. KORMAN. *Strahlentherap* 34, 508-604 (1929).—Blood catalase is regarded as storing radiant energy. B. C. A.

Micellar modifications of human serum by weak electrolytes. P. FOWLER. *Compt rend soc biol* 103, 1140-2 (1930), cf C. A. 24, 4813.—The effect of weak electrolytes on human serum is analogous to that obtained by strong electrolytes, acids or neutral salts. As the acidity increases, micellar vol., as measured by the viscosity of the serum, increases to a max. value and then diminishes. With an increased content of bile salts in serum, the micellar vol. as shown by decreased viscosity constantly decreases. The cond. of the serum increases constantly with increased acidity, but remains const. when the concn. of the bile salts in the serum varies. B. C. B.

The mechanism of the acidification of tissue culture medium. M. MAGATH. *Compt rend soc biol* 103, 1180-1 (1930).—The decrease in pH in tissue culture media from approx pH 8.0 to pH 7.0 is explained by glucolysis. CO_2 accumulates as a result of the action of lactic acid on bicarbonates. Carrel's method for regulating the pH of tissue culture by controlling the CO_2 content of the air around the tissues works satisfactorily. B. C. BAUNSTETTER.

Nucleic acid. I. Enzymes which split nucleic acid. YUTAKA JONO. *Ada Schol Med Univ Imp Kyoto* 13, 162-75 (1930).—The leaves and roots of green vegetables and seeds, in general, are richer in nucleic acid than are the fruits. A study of the P/N ratio shows that the base sugar compds. are more difficultly broken down than are the phosphoric acid sugar compds. Nucleic acid is rapidly attacked by plant nucleases, the hydrolysis reaching a max. in 3 to 4 days and continuing at an almost const. rate. Plant nucleases split yeast nucleic acid about 3 to 5 times as readily as they do thymus nucleic acid. The optimal pH for the reaction is 6.2 while a rapid decline in decomposition occurs in weakly alk. medium. The relationship between the setting free of purine bases and of phosphoric acid remains const. at different pH values. The nucleases are stable in acid or alk. medium and are more or less completely adsorbed by fuller's earth or kaolin. The reaction does not influence the completeness of this process. Animal nucleases were shown to be present in almost every organ but especially in the liver, pancreas, spleen, kidneys and small intestine while they occurred only in traces in the blood, muscles and the brain. II. The decomposition of nucleic acid through B. proteus. *Ibid* 176-81.—*Proteus vulgaris*, grown on agar media, was transferred to a sterile medium contg. nucleic acid and allowed to incubate for 7 days. The following

compds. were proved to be present in the bacterial digest of yeast nucleic acid either by m. p., elementary analyses or both: succinic acid, hypoxanthine, xanthine and uracil; cytosine was not found. With thymus nucleic acid the same products were obtained and also thymine. No cytosine was isolated from the animal nucleic acid III. The partial hydrolysis of nucleic acid through plant nuclease. *Id.* 182-6.—On the digestion of yeast nucleic acid with an ext. of soy-bean powder, guanosine, adenosine and cytidine were isolated but no uridine. These mols. were proved to contain pentoses. On digestion with an ext. of spinach, adenosine was also proved to be present. When thymus nucleic acid was treated with soy-bean nuclease, guanine hexoside was formed.

Denaturation of proteins by urea. W. RAMSDEN. *Nature* 126, 685 (1930).—The rate of denaturation of approx. isoelectric egg albumin by urea has a neg. temp. coeff.

The structure of the hemocyanins. I. The isolation of "hemocuprin," the copper component of hemocyanin. (*Octopus vulgaris*) ADOLF SCHMITZ. *Z. phys. & Chem.* 194, 232-47 (1931), cf. *C. A.* 24, 5771.—Hemocyanin, the blue pigment in the blood of molluscs and crustacea, may be coagulated without loss of respiratory function, and in this respect differs from hemoglobin. The moist coagulum retains the property of binding and again releasing O, as shown by alternate coloration and decolorization, but when completely dried it undergoes an alteration by which this property is destroyed. The Cu complex, however, is not split out unless the substance is treated with alkali. The linkage between metal component and protein is therefore more stable and of a different nature than that existing in hemoglobin. It is very resistant to acids, in fact more resistant than that of the metal itself, since the latter appears as ion before the entire complex is liberated. In the prepn. of the Cu component the finely powdered denatured hemocyanin is treated with *N* NaOH. The substance swells and soon takes on a reddish violet color. The color then gradually disappears but the soln. finally becomes dark and a green ppt. forms. The reaction is complete after 24 hrs. at 37°. After removal of the mother liquor which contains the protein component, the pptd. hemocuprin, as Na salt, is dissolved in hot H₂O or hot dil. EtOH to a dark wine-red soln. and reprecipd. by weak acids as a difficultly sol. green substance. Analysis gave the following values: C 45.04, H 6.99, N 12.55 and Cu 6.27%. P and S were absent. The substance has not yet been crystd.; hence the calcn. of an empirical formula was not attempted. The behavior of the substance toward solvents, the high H content which excludes an aromatic ring structure, and finally the spectroscopic behavior of the Na salt and the pyridine soln. of the acid show that a porphyrin structure is not present. Neither in structure nor in mode of linkage to protein does it resemble hematin. On the assumption that 1 mol. of hemocuprin contains 1 atom of Cu, the electrometric titration curve shows it to be a tetrabasic acid. It dissociates in 2 stages with dissociation consts. of approx. 4×10^{-7} and 8×10^{-10} , resp. Titration of hemocuprin explains the color changes which occur also with hemocyanin itself. The alk. soln., i. e., the quadrivalent ion of the dissohd. Na salt, is wine red. At *pH* 5.5 a development of blue color begins and increases up to *pH* 7.4, at which point the 2 weak acid groups have been neutralized, and the color is now pure blue. This represents the acid Na salt, i. e., the bivalent dissohd. ion. Further titration changes the color to green until finally the green undissohd. acid seps. out. The change of color to green, frequently observed with hemocyanin, has nothing to do with the formation of a hypothetical "methemocyanin"; it is a property of the undissohd. substance. Hemocuprin dissolves in concd. NH₄OH with a green color which changes to red when the soln. is diltd. It contains no free NH₃, but when hydrolyzed by 8 hrs.' boiling with concd. HCl practically all of the N becomes NH₃. Similar treatment with concd. NaOH converts only 24% of the N to NH₃. This behavior is the reverse of that usually observed with polypeptides. The Cu is evidently not linked to NH₃ groups.

Carbohydrate redoxase. HANS V. BULKE AND RAGNAR NILSSON. *Z. phys. & Chem.* 194, 290-8 (1931).—The hexosephosphate-dehydrogenase present in the seeds of yute (*Certhias capensis*) requires coenzyme for its activity. If the coenzyme of the seed ext. is removed by dialysis the enzyme action vanishes but can be restored by addn. of coenzyme. The instability of enzyme action in the ext. is not due to lability of the enzyme but to destruction of the coenzyme. Methylene blue reduction with fresh muscle and muscle ext. is activated by addn. of creatine. Washed muscle and dry preps., on the other hand, are not activated by creatine. Both types of muscle prepn. are activated by creatinephosphoric acid. It is possible, however, that the activation may be due to a small impurity in the creatine and creatinephosphoric acid preps. which exert a donor action.

A. W. DOX

Use of radium emanation and method of its administration. J. STRASBURGER, H. EAGLE
Klin Wochschr. 10, 29-31(1931)

Optical specificity of human liver esterase. P. RONA, R. AMMON, AND H. FISCHGOLD
Klin Wochschr. 10, 72(1931), cf. *C. A.* 25, 976—The results of Bamann and Laefferenz regarding the effect of strychnine upon the optical selectivity of human liver esterase in the hydrolysis of the racemic *Et* ester of mandelic acid have been confirmed. It has a similar effect upon the splitting of the *dl* methyl ester, facilitating the splitting of the *l*-ester. The greater velocity of the *l* hydrolysis is due to the fact that strychnine accelerates the splitting of the formed enzyme-*l*-ester combination by about 50%.
 H. EAGLE

The decomposition of the guanine nucleus by the enzymes of rabbit liver. GERHARD SCHMIDT
Klin Wochschr. 10, 165-7(1931)—A preliminary report. Details of the expts and results will appear in *Z. physiol. Chem.*
 H. EAGLE

Physiological chemistry as an independent profession. D. ACKERMANN.
Klin Wochschr. 10, 175-6(1931)
 H. EAGLE

Proteolytic enzymes of human white blood cells and blood serum. H. A. OELKERS AND H. FISCHGOLD
Klin Wochschr. 10, 205-7(1931)—Human white cells obtained by centrifugation from a blood specimen hemolyzed by acetic and tartaric acids contain erepsin, trypsin and cathepsin. Blood serum contains only erepsin.
 H. EAGLE

Biological-chemical methods in the investigation of phylogenetic problems. ROBERT JARETEKY
Arch. Pharm. 269, 50-62(1931)—An address.
 W. O. E.

Studies on pectin. V. The hydrolysis of pectin. ARTHUR G. NORMAN AND JOHN T. MARTIN
Biochem. J. 24, 649-40(1930), cf. *C. A.* 25, 1266—The yields of Ca pectate, furfuraldehyde and CO_2 were estd. at intervals in the course of pectin hydrolysis carried out at 100°, with 0.5% and 0.2% alkali. The rupture of the pectin ring proceeds more rapidly than the destruction of the furfuraldehyde-yielding groups. A polyuronide deriv. of pectin was prepd. by alk. hydrolysis of pectin.
 BENJAMIN HARKOW

Quantum problems in radiation biology. R. GLOCKER
Naturwissenschaften 19, 20(1931)—From the recent results of Wyckoff (*C. A.* 24, 5789, 25, 129) on the influence of radiation on *B. coli*, conclusions can be drawn on the basis of G's theory as to the effect of electron impulses in a cell. The dying of the bacteria follows an exponential curve for waves between 0.5 and 4 A. U., the exponent consisting of the radiation dose factor and a probability factor. From the latter is derived the "apparent vol." V of the cell which is sensitive for radiation. W. found that V changes with λ ; previous work of Lacaze and Holweck on *B. pyrocyanus* with $\lambda = 4-8$ A. U. (*C. A.* 23, 1928) gave a const. V . The mean path of an electron for $\lambda = 0.56-2.3$ A. U. is smaller than its range of action; hence the orbits of electrons originating outside V have to be included in the probability calcs. In accounting for this effect it is found that the true V is const. in all cases and that calcd. V values agree with the results of W. In the case of L and H the range of the electrons is far larger than their mean path and therefore does not affect V . For the calcs. the mean path inside the entire cell is to be used, not that inside the sensitive parts only ($1/10$ to $1/100$ of the total vol.). Apparently this indicates that several small sensitive spots are distributed all over the cell, a conclusion which is in agreement with the mol. conceptions of cells without nuclei. The change of V with λ and its const. value after correction for electron range is a direct proof for the quantum character of the influence of radiation on cells.
 B. J. C. VAN DER HOEVEN

The proteolytic action of papain and cathepsin. H. A. KAES.
Naturwissenschaften 19, 133(1931)—Recently Waldschmidt-Leitz (*Ibid.* 18, 952(1930)) denied an increase in proteolytic activity of papain and cathepsin by removal of heavy metal from their solns. and attributed the increase obtained by K. (*C. A.* 24, 3804) to the presence of HCN in the solns. freed from metal. In a further study it was found that at pH 5 HCN up to concns. of 0.2 mole has no effect at all. The action of complex forming substances (like HCN) depends on the nature of the metal impurity. If the metal is Zn as was proved formerly for this case, HCN has little effect; therefore removal of the metal is far more effective. If Au is present as inhibitor, HCN has considerable effect by changing the Au ions into complex ions (cf. also Graessmann in "Handbuch d. Biochem." of Oppenheimer, p. 184(1930)).
 B. J. C. VAN DER HOEVEN

Specificity of animal proteases. XXII. Mode of action of peptidases. ARNOLD K. BALLS AND FRANZ KÖHLER
Ber. 64B, 294-301(1931), cf. *C. A.* 25, 1542—In order to det. the mechanism of the reaction of aminopolypeptidase (I) of the intestine, the inhibitory effect of various substances on the action of I and of dipeptidase (II) was studied. The substances used were benzoylglycine (III), benzoyldiglycine, benzoyltriglycine, bromoisocaproylglycine (IV), bromoisocaproylidiglycine (V), bromo-

isocaproylglycine (VI), *p*-nitrobenzoylglycine (VII), aceturic acid (VIII), phthalimide (IX), sarcosine (X), allantoin (XI), creatinine (XII), glycine anhydride (XIII), benzoylglycine with changing concn of substrate, benzoylglycine with changing concn of enzyme, benzoylsarcosine and some amino acids (glycocoll, α -alanine, leucine, *d*-glutamic acid). From the results obtained certain conclusions may be drawn. That the inhibitory action of the NH group depends on its acid character is shown by the fact that III, IV, VII, VIII, IX and X inhibit the action of I; XI, XII and XIII do not. That only the peptide linkage adjacent to the acid residue has an inhibitory effect is shown by the fact that equiv quantities of IV, V and VI show the same inhibition of the action. Bromoisocaproylglycyltyrosine shows an increased inhibitory action, since it has more than one NH group to react with the enzyme. The cleavage of natural peptides may perhaps be explained thus—that the COOH activates the NH group for reaction with the enzyme. This view is strengthened by the fact that X, which has no free NH₂ group but an NH group, inhibits the action of I. Yet for other reasons it seems that the activation of the NH group is caused by the enzyme, *i. e.*, that the union of the enzyme with the NH₂ group of the peptides first renders the NH group capable of functioning as the second point of attachment. The results support the theory that the NH group functions as a point of attachment for all peptidases. This is the first common principle found for the action of peptidases. LOUISE KELLEY

A new proteolytic action of extracts of intestinal mucous membrane. ARNOLD K. BALLS and FRANZ KÖHLER. *Ber* 64B, 383-7(1931), cf preceding abstract and C A 25, 1542.—Although dipeptidase and aminopolypeptidase from intestinal mucous membrane attack the NH₂ and NH groups of peptides, glycerol exts of the membrane show a hydrolytic action of peptide-like substances which contain no NH₂ or COOH groups, *e. g.*, chloroacetyl-*o*-nitroaniline. The enzyme action described here belongs to a new type of peptidases which apparently react only with the NH group. The enzyme seems either to be present in very small quantities or to lose its activity easily as a result of inhibitory substances, for in the expts the increase in acidity was very slight. Expts with chloroacetyl-*p*-nitroaniline and benzoyltyrosine showed the optimum pH for cleavage to be about 8. LOUISE KELLEY

The crystallization, denaturation and flocculation of proteins with special reference to albumin and hemoglobin, together with an appendix on the physicochemical behavior of glycine. W. C. M. LEWIS. *Chem Reviews* 8, 81-165(1931).—Certain aspects of the behavior of denaturable proteins in respect to crystal, denaturation and flocculation are reviewed. A no. of closely related problems are also considered, among them the physicochem behavior of glycine in solution. LOUISE KELLEY

Action of sulphydryl, iron and cyanide compounds on the oxygen consumption of living cells. SANFORD M. ROSENTHAL. *U S Pub Health Repts* 46, 521-30(1931).—Glutathione, either in the oxidized or reduced form, when added in various tissues or to yeast cells, does not accelerate the rate of O consumption. Under the conditions of these expts, rat liver, brain, testicle, Jensen sarcoma and chicken erythrocytes are able to keep SH glutathione in the reduced state. The addn of considerable quantities of cysteine and Fe to these tissues does not overcome this property. Cysteine is also kept reduced by these tissues, even in the presence of added Fe. Rat kidney permits the slow oxidation of glutathione and cysteine. Blood serum causes a marked acceleration in the rate of oxidation of glutathione. Hematin, Na Fe tartrate or FeSO₄ (NH₄)₂SO₄ 6H₂O causes an acceleration of the O consumption of rat tissues or yeast cells. Methemoglobin, Na Fe tartrate and FeSO₄ (NH₄)₂SO₄ 6H₂O can prevent part of the inhibition of O uptake caused by NaCN. Oxyhemoglobin is less effective, while hemin shows no effect. Neither oxidized nor reduced glutathione influences the inhibition of respiration by cyanide. However, a slight antagonism can be demonstrated if S-S glutathione and cyanide are allowed to react for a considerable time before they are added to the tissue. KCN and α -amino- β -sulfolopropionic acid do not inhibit the O consumption of rat testes. KCNS causes a slight increase in the rate of O consumption of rat testes. The significance of these results with respect to the mechanism of the action of cyanide upon living cells is discussed. J. A. KENNEDY

Kinetics of esterase activity in comparison with acid catalysis. ERNST A. SYM. *Biochem Z.* 230, 19-50(1931).—To avoid inactivation of esterase by the substrate and to provide a suitable solvent for the various org. acids, alcs and esters, an acetone medium was found most favorable. In studying the inactivation of the esterase in the presence of 0.2 mol AcOH per l it was found that this becomes less as the mol wt. of the alc. increases (up to butyl alc) and this is also true for the homologous fatty acids of the aliphatic series. The inactivation of the esterase by H₂O in an acetone medium becomes considerable at a concn. of 6-8 mols. per l. Comparative expts.

with esterase and with HCl catalysis at different concns of butyl alc. showed that the reaction velocity const. is subject to large variations. Water activates esterase and inhibits HCl catalysis. The esterase activity increases with a rise in the butyric acid concn, very rapidly up to a concn of 0.0 mol per l; then it remains const or even diminishes. The velocity of the esterification reaction of different alcs of the homologous series, with esterase or HCl, or under the influence of heat is practically the same. However, the influence of homologous acids on the kinetics of esterification is entirely different for esterase than for the HCl catalysis or for the thermal reaction. The hypothesis is developed of the ester formation by a process of contact catalysis.

S MORGULIS

Characterization of proteins through the determination of their affinities. G. ERTSCH, H. SACHSSE AND W. BUCK. *Biochem Z* 230, 68-92(1931).—The affinity of the proteins for Cu is measured by the e. m. f. of the compd., and this is offered as a means of identifying the different proteins. The reaction between alk. protein solns and CuSO_4 (biuret reaction) has been studied quantitatively, the e. m. f. being detd. in relation to the protein, Cu, alk. concns and temp in the following chain: $\text{Cu} \mid \text{protein} + \text{NaOH} \mid \text{KCl} \mid \text{CuSO}_4 \mid \text{Cu}$. The influence of the alkali at the end of 24 hrs.

+
 CuSO_4

reaches a definite point which could be detd. not only by studying the e. m. f. but also changes in depolarization, viscosity and light absorption. Electrochem. differences become apparent especially in high protein concns, when the e. m. f. with rising NaOH concn is found to increase much more with albumin than with globulin. On the contrary, with increasing Cu concn, the e. m. f. of the albumin diminishes more rapidly than that of the globulin. These facts are interpreted as showing that in albumin the Cu binding groups have a greater affinity for the Cu than in globulin, but that their no. is smaller in albumin than in globulin. Certain theoretical and methodical points are discussed in connection with this study.

S MORGULIS

Depolarization and light absorption of alkaline protein solutions. G. ERTSCH, H. SACHSSE AND D. LANGE. *Biochem Z* 230, 93-114(1931).—The depolarization of alk. protein solns varies with time and with the alkali concn. These alterations are due to the swelling as well as breaking up of the protein particles, which only occur in the presence of sufficiently large alkali concns. Globulin is more easily broken up than albumin.

S MORGULIS

Viscosity of alkaline protein solutions. G. ERTSCH AND H. SACHSSE. *Biochem Z* 230, 115-28(1931).—The flow elasticity of albumin or globulin solns. shows no measurable alteration within a wide reaction range, but there is an indication of a form elasticity as soon as a few cc of CuSO_4 soln is added to a highly concd. alk. soln of albumin or globulin. In fact under some conditions the solns. may become solidified on the addn. of the CuSO_4 . The viscosity increases with increased alkali addn. after a certain concn (satn. concn) has been exceeded which causes far reaching alterations in the protein mol. The ability to split up is greater in globulin than in albumin, and this is considered to be responsible for the fact that the globulin is more hydrophile.

S MORGULIS

The chemical independence of the serum proteins. G. ERTSCH AND H. SACHSSE. *Biochem Z* 230, 129-35(1931).—The euglobulin fraction obtained by 1% satn. with $(\text{NH}_4)_2\text{SO}_4$ corresponds to the globulin fraction by electrodialysis. The paraglobulin obtained by half satn. with $(\text{NH}_4)_2\text{SO}_4$ is an independent chem. substance, which occupies a definite position between the globulin and albumin. The transformation of one into another is not very probable.

S MORGULIS

Influence of lecithin on the stability of the serum proteins. ISTVÁN WENT AND FERENC FARAGÓ. *Biochem Z* 230, 238-44(1931).—On mixing an aq. lecithin emulsion with serum a definite reciprocal relationship is manifested in the alteration in the amts. of euglobulin and pseudoglobulin I on the one hand, and between pseudoglobulin II and albumin, on the other. Conclusion: Euglobulin is a colloidal complex of pseudoglobulin and lecithin, and the same may also be true for the pseudoglobulin II (pptd. by 21.5% Na_2SO_4) and albumin. Between the 2 main types of serum protein there seem to be less definable dispersoids whose precipitability by neutral salts varies greatly according to the amt. of lecithin.

S MORGULIS

The specificity of phosphatase. KURT P. JACOBSON AND JOAO TAPAOINHAS. *Biochem Z* 230, 304-11(1931).—Mono- and dimethyl phosphate and dimethyl pyrophosphate esters could not be hydrolyzed by phosphatase.

S MORGULIS

The destruction of cystine and of cysteine through alumination. FRITZ LIEBEN AND ERICH MOLNAR. *Biochem. Z.* 230, 347-52(1931).—Both cystine and cysteine

undergo decomposition under the influence of light from the quartz Hg lamp or diffuse daylight, but this proceeds at a much slower rate than with aromatic amino acids. In diffuse daylight the destruction of cystine goes on only in an alk. medium, while that of cysteine must be followed in an acid medium since it oxidizes spontaneously in an alk. medium. Hematoporphyrin must be used as a sensitizer with the cysteine. Under the influence of the illumination with the quartz Hg lamp no sensitizer is necessary, but the rate of the reaction without it is markedly slower. The products of radiation of cystine when added to tryptophan or to tyrosine diminish their color reactions. The reduction of methylene blue in diffuse daylight by cysteine as compared to that in the dark shows that with an increase in the acidity the difference in time required for the fading of the color becomes constantly greater while at the same time the rate of fading sharply decreases.

S. MORGULIS

Refractometric studies on serum protein. III. The specific refraction for the total protein and for the non-protein substances of horse serum. D. VON DESEB. *Biochem. Z.* 230, 373-82(1931), cf. *C. A.* 24, 1874.—The refractive index and sp. gr. were detd. at 17.5°, also the total protein gravimetrically, in 30 horse serums, both normal and immune. A study of the data leads to the following formulation. The refractive index of the serum, $R_1 = 1.33320 + 0.00209 + 0.00187 \times \text{protein content}$, with an av. error of ± 0.00002 . Or, specific gravity of the serum, $S.G. - 1.0 = 0.00732 + 0.00268 \times \text{protein content}$, with an av. error of ± 0.00003 . The R_1 and $S.G.$ of the non-protein materials of the serum are 0.00209 and 0.00268, resp. The increase in the specific refraction of serum protein is 0.00187 and that of the specific wt. 0.00268 per unit of protein. IV. Increase in the specific refraction of the protein fractions in horse serum. *Ibid.* 383-94.—Three groups of serum proteins were studied from the point of view of their n . The first group was represented by the protein pptd. by satn. with a definite amt. of $(NH_4)_2SO_4$, and the second by those remaining in soln. in the serum. The former comprised all the globulins with an av. n of $0.00195 = 0.00001$ and a corresponding sp. gr. of $0.00298 = 0.00005$. Corrected on the basis of sp. gr. data, the n should be 0.00187, which is the same as that for the albumin. The third group was represented by mixts. of both albumins and globulins. Here the n was $0.00185 = 0.00002$ and the sp. gr. $0.00272 = 0.00005$. It has been further substantiated that for 1% serum protein the variations in n and sp. gr. are, resp. $0.00187 = 0.00001$ and $0.00268 = 0.00005$.

S. MORGULIS

Hydrotropic solution of calcium, with reference to the solution of calcium in blood serum. A. VON KUTHY AND H. BANGA. *Biochem. Z.* 230, 458-65(1931).—Ca dissolved hydrotropically in a Na salicylate soln. behaves like a serum Ca in that part of it is ionized, is negatively charged and part is in a non-diffusible condition. It is therefore suggested that serum Ca is hydrotropically dissolved.

S. MORGULIS

Comment on the paper of Schreiber and Friedrich: "Demonstration and intensity of mutagenic radiations." A. GURWITSCH. *Biochem. Z.* 230, 503(1931).—Discussing the neg. findings of these authors (*C. A.* 25, 529) G. points out that agar yeast cultures do not radiate in the dark but can still be used as detectors.

S. MORGULIS

Investigations on the source of oxidase granules. GENTER WOLBACH. *Folia Histochem.* 43, 121-31(1931).—The subcutaneous injection of horse-radish or beet ext. produces oxidase granulation in white mice.

JOHN T. MYERS

The dependence of oxidizing fermentations upon the oxidation reduction potential of the external medium. S. I. KUZNETZOV. *Zentr. Bakteriol. Parasitenk.* 2 Abt., 83, 37-52(1931).

JOHN T. MYERS

The enzymic hydrolysis of gelatin in its relation to the formation of discipiperazines (BLANCHETIERRE) 10. Neo-xantho-neobilirubin acid and partial synthesis of meso-bilirubin and mesobilirubinogen (FISCHER, HESS) 10.

BRault, A. Le glycogène dans le développement des tumeurs, des tissus normaux et des êtres organisés. Paris: Masson et Cie. 370 pp. F. 80.

GENEVORS, L.: Métabolisme et fonction des cellules. Paris: Masson et Cie. 118 pp. F. 26.

PARKSONS, THOMAS R.: Fundamentals of Biochemistry in Relation to Human Physiology. 3rd ed., revised. Cambridge, Eng.: W. Heffer & Sons, Ltd. 308 pp.

VERNE, JEAN: Couleurs et pigments des êtres vivants. Paris: Armand Colin. 219 pp. F. 10 50. Reviewed in *Nature* 127, 365(1931).

B—METHODS AND APPARATUS

STANLEY R. BENEDICT

Determination of p_n of [biological] media. M. LEMOIGNE AND R. CHAMINADE. *Compt rend soc biol* 102, 922-3(1929)—When the colorimetric method is inapplicable and the electrometric method troublesome, the medium may be brought into electrometric equilibrium with a phosphate mixt. of colorimetrically detd p_n . B. C. A.

Detection of necrosis of pancreatic tissue by determination of urinary diastase. J. WOHLGEMUTH. *Klin Wochschr*, 8, 1253-4(1929)—W.'s method for the detn of urinary diastase in presence of a phosphate mixt of p_n 7.2 is described. B. C. A.

Microdetermination of proteins. R. WINTERNITZ AND Z. STARY. *Mikrochemie* [N. S.] 2, 252-6(1930)—A quantity of the soln contg. not more than 15 mg. of protein is centrifuged for 15 min. with 2 cc. of 20% trichloroacetic acid, the liquid decanted, and the pptd. washed with 4% trichloroacetic acid, and heated with a mixt. of H_2SO_4 and H_3PO_4 to destroy org. matter. The mixt. is dild. to 17.5 cc., 7.5 cc. of Nessler soln is added, and the color developed is compared with a series of standards. B. C. A.

A critique of present methods for the study of gastric acidity. MILTON J. MATINEA AND IRVING GRAY. *Arch. Internal Med* 47, 53-63(1931)—The methods of gastric analysis are discussed. Histamine should be used as a gastric stimulant if the routine test gives no free HCl. A definite alk. tide was found in 50% of the cases showing free HCl after injection of histamine. J. B. BROWN.

The detection and estimation of lactose in urine. C. P. STEWART AND R. M. GRAY. *Edinburgh Med J*, 38, 109-12(1931)—The total sugar (A) in the urine is detd. by Cole's method and calcd. as glucose. The total sugar (B) in urine hydrolyzed with HCl is calcd. in the same way. The amt. of glucose (y) and the amt. of lactose (x) can be calcd. from the equations $y + x/2 = A$, $y + x = B$. The method is suitable for the examn. of urine during pregnancy and lactation. R. B.

The technic of tissue culture in hanging drops. ALEXIS CARREL. *Compt rend. soc biol* 102, 742-4(1929)—The p_n of tissue cultures is regulated by flooding a culture chamber with expiratory breaths until an indicator drop (phenol red in a drop of medium) shows that the desired p_n has been reached. The chamber is formed by a metal ring, 5 cm. in internal diam. and 1 cm. high; at each end of the ring is a ledge, 0.25 cm. deep. A mica sheet is affixed to each end by paraffin. At one end of the ring there are 2 openings to opposite walls, provision is made for the insertion of a metal cannula, to which rubber tubing may be attached, to flood the chamber with expired air. Four tissue cultures may be observed simultaneously. B. C. BAUNSTETTER.

A diazo method for detecting bilirubin in urine. GEORGE HUNTER. *Can Med Assoc J* 23, 823-4(1930)—The possibilities of detecting bilirubin in icteric urine have been reinvestigated, and a satisfactory qual. test, which might form the basis of a quant. method, is as follows. To 5 cc. of urine in a 15-cc. centrifuge tube add 2 cc. of 10% $BaCl_2$ mixt. and centrifuge. Pour off the supernatant fluid, wash the ppt. with a few cc. water, centrifuge again and pour off the liquid. Add 0.5 cc. diazo reagent, stir the ppt. with a glass rod, and add 2 cc. of 96% alc., and, as buffer, 0.3 cc. of 6% Na_2HPO_4 , 12 H_2O . If bilirubin is present the same color results as with acid. If it is highly pigmented, alter the Ba ppt. is obtained stir with 2 cc. of 96% alc. to which has been added 1 drop of 10% H_3PO_4 , centrifuge and transfer the supernatant fluid to a second tube along with 4 cc. water. Add 2 drops 10% $(NH_4)_2SO_4$ and 2 cc. $BaCl_2$ ceptg. the bilirubin, and proceed as above. A. T. CAMERON.

The colorimetric determination of nonprotein residual nitrogen in blood. G. A. BROSSA. *Atti accad. sci. Torino (Classe di sci. fis., mat., e nat.)* 65, 272-4(1930)—Methods of analysis which may be applied readily without very elaborate app. or technic are being developed in biochemistry as an aid in diagnosis (e. g., tests for sugar, uric acid, etc.). A method is described for detg. nonprotein N without the necessity of any distn. Dealbuminize 1.2 cc. of blood serum to 8.8 cc. distd. water by adding 4 cc. of 10% sulfosalicylic acid soln. After 30-60 min. filter the soln., and, after adding 1 cc. of H_2SO_4 , and a crystal of K_2SO_4 , evaporate it to dryness. Redissolve the residue with water to exactly 12 cc. Use 2 cc. to det. the amt. of 30% NaOH necessary to make the soln. alk. to litmus. Then add the calcd. amt. to the 10 cc. left and add 2 cc. 10% PbOH soln. and 0.6 cc. of freshly prepd. NaOCl soln. Heat this mixt. as well as a standard soln. for comparison in boiling water for 10 min., cool and compare the colors with the help of a colorimeter. A. W. CONTIERI.

The calculation of the color index of blood. FRANKISKA STENGEL. *Hesper*

Klin. Wochschr. 44, 194-7(1931).—An alignment chart shows color index, in terms of erythrocyte count and quantity of hemoglobin. The latter is detd. by the Hellige or the Sahli method. D. B. DILL

Determination of small quantities of zinc in the presence of lead salts. M. W. STAS. *Pharm. Weekblad* 68, 93-7(1931).—Zn may be detd. by pptn with *o*-hydroxyquinoline in solns. contg. 1% AcOH and 4% NaOAc in the presence of Pb, provided the concn of Pb does not exceed 100 mg. per 50 cc. The pptd. Zn salt of *o*-hydroxyquinoline is detd. volumetrically by bromination with KBrO_3 + KBr and titrating the excess Br with $\text{Na}_2\text{S}_2\text{O}_3$ as recommended by Berg (*C. A.* 21, 2630). A. W. DOX

Titration of organic acids in the urine. WILHELM WERTZ. *Klin. Wochschr.* 9, 1632(1930).—Occasionally the tint of the urine precludes or makes difficult the use of tropeolin OO as indicator in the titration of urine org. acids by the Van Slyke-Palmer technic. In such cases W. recommends the use of β -dimethoxyphenol. H. EAGLE

New reactions of bilirubin in blood serum. GABRIEL MONASTERIO. *Klin. Wochschr.* 9, 1772-3(1930).—If one adds 1 drop of H_2O_2 and 1 drop of glacial acetic acid to 5-6 drops of serum, a green color develops in the presence of bilirubin, which is intensified by heat. By heating serum with the diazo reagent of van den Bergh, sera which would otherwise give an indirect reaction give a direct reaction. Alc. is a similar "catalyst." H. EAGLE

Microestimation of the nitrogenous constituents of urine. LUDWIG PRINCUSSEN. *Klin. Wochschr.* 9, 1966-7(1930).—A reply to the criticism of Jacoby (*C. A.* 24, 5775). Princussen believes 15 min. distn. quite adequate. H. EAGLE

Estimation of blood volume in congenital heart defect. KARL HITZENBERGER AND FRITZ TUCHELD. *Klin. Wochschr.* 9, 2159-60(1930). H. EAGLE

Hormonal reaction for pregnancy with the urine of humans and animals. BEN-HARD ZONDER. *Klin. Wochschr.* 9, 2255-9(1930).—The hypophyseal hormone has been demonstrated in the urine of pregnant women, apes and horses, but could not be found in cows, pigs, elephants or Rodentia. The urine of pregnant mares contains about 10 times as much folliculin per unit vol. as that of pregnant women, but, unlike the latter, is in a form not sol. in org. solvents. From such urine it is possible to prep. large quantities of cryst. hormone, contg. 8000 units per g. The folliculin content of blood is only $1/100$ that of the urine. In marked contrast, the blood of pregnant horses contains large quantities of hypophyseal anterior lobe hormone, as shown by the ovarian hyperplasia and the formation of corpora lutea induced by its injection into infantile rats; but the urine contains only minimal quantities. The longer the pregnancy, the less of the hypophyseal hormone circulates in the blood. Z. distinguishes 2 types of the hypophyseal hormone: (1) HVH-A, which causes ripening of follicles, and (2) HVH-B, which causes luteinization. No explanation is offered for the massive excretion of folliculin by the pregnant horse, and the retention of the HVH, particularly HVH-B. It is important to note that in the diagnosis of pregnancy in women, the essential criterion is the presence of HVH-B in the urine, which induces luteinization and the formation of blood spots on the ovary of the test animal; both folliculin and HVH-A may appear in the urine during other conditions, such as at the menopause, in amenorrhea, etc. In the horse, however, the best criterion is the induction of rut in the infantile rat, as shown by the appearance of cells in the vaginal smear. This is a test for the enormous increase in the urine folliculin and is more accurate if the urine to be tested is first acidified, filtered and extd. with ether to remove an uncharacterized inhibiting factor. In 9 virgin mares, this test was uniformly neg.; in 54 pregnant, it was positive in 53; and in 17 mares which had been mounted but had not become pregnant, the test was pos. only once: a total error of only $2\frac{1}{4}\%$. H. EAGLE

A new method for the estimation of glucuronic acid in the urine. JOACHIM SAUER. *Klin. Wochschr.* 9, 2350-1(1930).—The ether ext. of the acidified urine is evapd. to dryness, heated with 12% HCl, and the glucuronic acid estd. by the usual Lefèvre app. The av. daily excretion is 0.22-0.29 g. H. EAGLE

A method for the determination of organic iodine (uroselectan) in the urine. K. HILLGRUBER. *Klin. Wochschr.* 9, 2353-4(1930).—The urine is oxidized with hot H_2SO_4 and H_2O_2 ; the I, now free from its org. radical, is then oxidized with Cl water in alk. soln. to HIO_3 , and this is titrated in acid reaction with KI and $\text{Na}_2\text{S}_2\text{O}_3$. H. EAGLE

The peroxidase reaction. XXVIII. An exceedingly sensitive peroxidase reagent for human milk. TAYURU ARAKAWA. *Tôhoku J. Exptl. Med.* 16, 83-9(1930).—Reagent A consists of guaiac resin 0.3-1 g., As_2O_3 0.02 g., glacial AcOH 0.6 g., AcONa 1.30 g., and 99% EtOH to make 100 cc. Reagent C consists of pharmacopeial tincture with 0.1% peroxide (cf. *C. A.* 25, 1330) I, and guaiacol 2 parts, and acetone to make

100 cc. A mixt. of *A* and *C* is used for qual. tests. For quant. tests reagent *B* (benzidine 1 g., AcONa 1.26 g., and 99% alc. to 100 cc.) and reagent *D* (described in part XXX, below) are used. Despite the fact that the qual. test is 64 times as sensitive as those currently used, it fails to react with blood and can, therefore, be used for testing pus, spinal fluid and milk which are contaminated with blood. **XXIX.** Determination of human milk peroxidase. 1. Dilution method. *Ibid* 97-98—A d.in. method in which are used reagents *A* and *C* (part XXVIII, above) is described. **XXX.** Determination of human milk peroxidase. 2. Colorimetric method. *Ibid* 97-106—The reagents used are (1) reagent *B* (part XXVIII, above), (2) reagent *D* (a 3% soln. of H_2O_2 0.1, and guaiacol 2 parts and 99% EtOH to make 100 cc.), (3) Walpole's acetate mixt. at pH 3.6 dild. 1:10 with 0.9% NaCl , (4) CHCl_3 , (5) 0.1 *M* AcONa in 99% alc., (6) a standard soln. consisting of a CHCl_3 ext. of $\text{Cu}(\text{OAc})_2$. Equal vols. of (1) and (2) are added to one vol. of milk (whole, or d.M. with (3)), and the mixt. is extd. with CHCl_3 . The ext. is made up to 5 cc. with 99% alc., and compared colorimetrically with the standard. **XXXI.** An approximate estimation of milk peroxidase. *Ibid* 107-11—To 1 cc. of milk are added 3 cc. of *M* 15 phosphate buffer at pH 7.8, 1 cc. of reagent *A* (above) and 1 cc. of reagent *C* (above). In a milk contg. large quantities of peroxidase, a deep-blue color develops immediately. **XXXV.** A rapid macro method for peroxide by the use of milk peroxidase; preservation of milk peroxidase. *Ibid* 232-5—To 0.5 cc. of reagent *A* are added decreasing quantities of the substance to be tested, and 0.25 cc. of standard peroxidase; the color which develops is compared with that formed by a known peroxide soln. (3% H_2O_2 dild. 1:100). Milk peroxidase may be preserved indefinitely by the following technique: 100 cc. of milk is shaken with 10 cc. CHCl_3 , 10 cc. 95% EtOH and 1 cc. guaiacol, 0.5 cc. glacial AcOH is added and the mixt. filtered 24 hrs. later. The filtrate keeps indefinitely. H. EAGLE

Sulfosalicylic acid as protein reagent. HUGO KCM. *Süddeutsche Apoth.-Ztg.* 70, 665-6(1930).—Sulfosalicylic acid is a useful reagent in physiol. chemistry, notably in the detection of albumin in urine. In the present study expts. have been carried out to characterize the various ppts. obtained by treating certain sol. proteins with 1 cc. of the reagent. The protein solns. examd. contained 0.5% albumin. Among the proteins examd. were albumin Merck and Sanguin, natural albumin of duck eggs, meat albumin, plant albumin, gelatin, peptone, keratin, etc. On addn. of the reagent the soln. is heated to boiling, and observations are made on the time required to produce a ppt. (if any), the character of this ppt., being noted. The value of sulfosalicylic acid as a precipitant lies in the fact that the nature of the ppt. often enables one to determine what protein is involved. W. O. E.

Chemical detection of blood according to Segeloff. PAUL SCHROTT. *Pharm. Presse, Wiss.-prakt. Hefte* 1930, 166-7.—This test, applicable to the detection of blood in the urine or feces, in conjunction with spectroscopic and microscopic tests, involves the following procedure: While the urine may be tested directly, the feces require a preliminary treatment, as titration of a 20-g. sample in a mortar with a like amt. of 0.5% AcOH , followed by filtration through a wet (if necessary, double) filter. To 5 cc. 5% pyrazinone soln. in abs. EtOH, add 8 drops 3% H_2O_2 soln., 8 drops 50% AcOH and 10 drops of the urine or feces filtrate. The presence of the minutest traces of blood in either sample is shown by a violet to light-blue coloration. This test is about 10 times more delicate than that of Alimen and about 3 times more sensitive than the Adler test. An improvement on the Segeloff procedure consists in treating 1 cc. of the sample in a test tube with 8 drops AcOH and 8 drops H_2O_2 , and then, with the tube in an oblique position, adding, drop by drop, the pyrazinone reagent. The first 10 or 15 drops mix with the soln. mixt., whereas further addn. of the reagent develops a layer superimposed on the lower, the zone of contact showing a light-blue to amethyst-violet coloration. W. O. EMERY

Selective staining of basophilic granules. EUGENE BUJARD. *Bull. histol. appl.* 7, 284-9; *Sém. Tech.* 6, 31(1931).—The technique is described. C. R. FELLERS

Histochemical detection of hemoglobin. L. LEROY. *Rev. hyg. med. pr. S.* 894-6, *Compt. rend. soc. biol.* 103, 36-8(1930).—Hemoglobin in blood is transformed to hematin by treatment with the following mixt.: Pb acetate 2-5 g., CH_3O 10 cc. and distd. water 100 g. Peroxidases in the hemoglobin are annihilated by the treatment and give a characteristic benzidine H_2O_2 reaction (blue color). The reaction is sp. C. R. FELLERS

Colorimetric determination of phenols in feces. DEBERGH AND R. GOUDON. *Per. hyg. med. pr. S.* 53, 78-80(1931).—Dil. 10 g. of feces to 100 cc. and treat 5 cc. of this soln. with 5 cc. of 10% Na tungstate and 5 cc. of 0.66 *N* soln. of HCl . In a few min. stir the soln. and filter. To 6 cc. of the filtrate, corresponding to 2 cc. of the 1:10 diln.

of feces, add 5 cc. of an aq. soln. of 2.5% ZnCl_2 and 5 cc. of 20% Na_2CO_3 soln. Agitate the mixt. and filter. Place 8 cc. of the clear filtrate in a 30 cc. cylinder together with the phenol reagent and make up to 25 cc. with distd. water. Finally add 5 cc. of a 20% soln. of Na_2CO_3 , shake the mixt. and set aside for observation at the end of 1 hr. The development of a blue color indicates the presence of phenols. The amt. may be estd. by colorimetric comparison with a standard made up by adding 0.001 g. of resorcinol per l. and treating 1 cc. of this soln. by the method just described. This amt. of resorcinol gives the same intensity of color as 0.1 g. of phenol per l. The method is considerably more accurate than the original Fohn and Denis technic. C. R. F.

New rapid method for tissue diagnosis. CHARLES F. GRESCHICKTER, EDWARD P. WALKER, A. M. HJORT AND CARL H. MOULTON. *Stain Tech.* 6, 3-12(1931).—A new staining procedure applicable to both fresh and CH_2O -fixed frozen sections is described. A pre staining bath consisting of KH_2PO_4 0.75 g., N NaOH 30 cc., distd. water 570 cc., glycerol 200 cc. and 95% EtOH 200 cc. is used to collect the sections as they come from the microtome. The sections are then passed into the following staining soln. for 20-30 sec. thionine eosinate 0.75 g., Ba eosinate 0.25 g. and azure A 0.25 g. dissolved in 100 cc. of a mixt. of 4 parts of (Cl_2O) , and 1 part of EtOH 95% plus 0.2% glacial AcOH . The stain is washed off in 2 successive changes of 95% EtOH contg. 200 cc. of glycerol to each 80 cc. of EtOH . The tissue is left in the first bath about 10 sec. and in the second bath 3 sec. Sections are dehydrated by floating for 10-15 sec. in a soln. of diethylene glycol monobutyl ether and cleared in butyl phthalate by floating in this soln. for 20 sec. The section is finally floated on a glass slide, blotted dry with photographic lintless blotting paper and covered with an inverted cover slip upon which 1-2 drops of dammar gum has been placed. C. R. F.

Dissection, staining and mounting of styles in the study of pollen-tube distribution. JOHN BUCINOLZ. *Stain Tech.* 6, 13-24(1931).—Details of the method are given.

C. R. FALLERS

Method of staining pollen tubes within the pistil. CLYDE CHANDLER. *Stain Tech.* 6, 25-6(1931).—A very satisfactory killing soln. consists of 6-7 cc. of conc. CH_3O in 70% EtOH . The method of dissection and staining with acetocarmine is outlined.

C. R. FALLERS

Lacmoid-Martius—yellow for staining pollen tubes in the style. B. R. NANEL. *Stain Tech.* 6, 27-9(1931).—The staining technic is described. The callose stain permits of interpreting the physiol. condition of the tube more readily than a stain of the plasmatic content alone. The method is particularly valuable with fleshy styles such as are found in pomaceous plants, cherries, plums and grapes. C. R. FALLERS

Method for detecting mineral particles in tissues, particularly in lung tissue. A. POLICARD. *Bull. hist. appl. physiol. et path.* 7, 120-31(1930); *Physiol. Abstracts* 15, 203; cf. C. A. 24, 3579.—A method is described for detecting such particles in lung tissue, but it is applicable to other investigations. Conc. HClO_4 is used to dissolve or render transparent the org. particles. The paraffin section is treated with xylene, then alc., and dried; it is covered with a drop of strong HClO_4 and heated on a boiling water bath until it becomes transparent; it is immediately cooled and examd. The method is of value in studying the mineral deposits found in the lungs in certain miners' diseases. G. G.

Ultrafiltration in vivo. L. BAULL. *Compt. rend. soc. biol.* 99, 1605-7(1928); *Physiol. Abstracts* 14, 482.—The principle of the ultrafilter of Giemsa is adapted to an app. to receive the blood from the carotid of a dog after the injection of heparin to prevent coagulation. The ultrafiltrate from the plasma is obtained in sufficient quantity to permit physicochem. study. G. G.

Action of ox bile in filtration. D. COMPRESO AND A. DAMBOVICRANU. *Compt. rend. soc. biol.* 103, 182-4(1930); *Physiol. Abstracts* 15, 80.—Ox bile was used to facilitate the passage of substances in emulsion through filters, Chamberland bougies and collodion sacs. Victoria blue and Congo red do not pass in ordinary conditions, but the latter in 1 per 2,000 soln. passes through the bougie to a slight extent. Addn. of bile (1 in 10 to 1 in 100) facilitates the passage of both. *Bacillus pyocyaneus* also passes through the bougie after the use of bile. G. G.

Colorimetric estimation of phenols in feces. NME. TONCKIHERE-DEBRAGH AND R. GOUFFON. *Compt. rend. soc. biol.* 103, 435-7(1930); *Physiol. Abstracts* 15, 207.—Errors occur in the filtration of tungstic-molybdic ppt. in the method of Fohn and Denis. These are obviated in the method employed by the authors for detg. phenols in feces. In this process the tungstic filtrate is treated with ZnCl_2 and Na_2CO_3 . The ZnCO_3 retains the S, which is capable of acting on the colorimetric agent, and the resulting color is easily compared with the standard resorcinol color. G. G.

Importance of protein precipitation in amino acid determination in blood. P. M. RE AND D. POTICK. *Rev. soc. argentina biol.* 5, 725-44 (1929).—Folin's method gives values below the real ones, and the error increases with the concn. in amino acids. The tungstate H_2SO_4 mixt. used to ppt. the proteins diminishes the values obtained up to 50%. Trichloroacetic acid gives a much smaller error. Phosphotungstic acid gives intermediate values. G. G.

Methods for the study of the physiology and pharmacology of the artificially perfused mammalian intestine. H. F. ROESSA. *Arch. ges. Physiol. (Pflüger)* 226, 171-83 (1930).—Various methods are described for perfusing mammalian intestine. ARTHUR GROLLMAN

Microchemical examination of very small quantities of skin for phenolic substances. HANS SCHMALFLOSS AND HELENE BARTHMEYER. *Mikrochemie* (N. S.) 2, 245-52 (1930).—*Melanogen* is a provisional name given to any substance which turns an enzyme strip of the flour beetle, *Tenebrio molitor* L., brown, gray or black. In studies concerning inheritance and development, it was desirable to test for the distribution of melanogens, e. g., α -dihydroxybenzenes, in organisms, particularly in the colored skins of insects. By means of simple methods, most of them already known, it was found possible to do this with very small quantities of material. To test the color, $1-6 \times 10^{-4}$ cm. of material was pulverized, a suspension made with 0.008-0.25 cc. of water, the suspension was transferred to a small test tube and heated for 2 min. in boiling water to destroy injurious enzymes. The soln. was then filtered and tested with enzyme test strips. Tests were also made with FeCl_3 - Na_2CO_3 for dihydroxybenzenes and for phenols by Gibb's reagent and with Millon's reagent. The work is described in detail and will be interesting to those who desire to make similar tests. W. T. H.

A modification of MacCallum's hematoxylin method for iron. R. R. DIETHELM. *Arch. Path.* 10, 740-1 (1930).—In the original method of MacCallum for the staining of unmasked Fe a freshly prep'd 0.5% aq. soln. of hematoxylin was used. The modification consists in the addn. of 1 cc. neutral CH_3O soln. to 100 cc. of the hematoxylin soln. The CH_3O acts as a reducing agent and prevents the natural oxidation of the hematoxylin soln. Such a CH_3O hematoxylin soln. has been used for 3 months without losing its ability to combine with the Fe. HARRIET P. HOLMES

The state of acidity of the stomach contents and its clinical estimation. F. EBERV. *Wiener Arch. inn. Med.* 20, 353-74 (1930), cf. *C. A.* 24, 3527.—By the use of the phenolphthalein test of Gunzberg, the free HCl can be det'd. in the stomach contents. This is no test for H ion content as with the same pH value, solns. of HCl, tartaric acid and citric acid react positively and solns. of normal fatty acids and lactic acids negatively. The HCl bound to proteins cannot be det'd. by any direct titration method. Its estn. follows the detn. of free HCl and of the total HCl after the methods of Sjögqvist or Lüttke-Martius. By use of these methods it is possible to distinguish in the stomach contents between HCl as the normal acid of the stomach and the accidental acids of the food or pathologic org. acids. The detn. of the actual acidity directly or indirectly (titration of the indicator according to Sahli) is not superfluous or to be ignored but is supplemented. As the taste of the acids is not due to the H ion concn. alone, but to the equiv. concn., the equiv. concn. must be considered as well as the actual acidity. In order that results may be compared in writing on the acidity of the stomach contents the nature of the test meal should be given and the methods used, such as total acidity (α -naphtholphenathalein) or total acidity (phenolphthalein), free HCl (dimethyl yellow), actual acidity (Sahli). HARRIET P. HOLMES

A new simple method for preserving microscopical stained films. M. VAN RIEMSDYK. *Nederland. Tijdschr. Hyg. Microbiol. en Serol.* 5, 119-24 (1931).—An aq. soln. of transparent gum arabic can be applied with success as a film cover to all microscopical stained smears, perhaps even to histological and anatomical films. The gum recommended is the Gee Wee gum. Take 1 part of gum and 1.5-2 parts of H_2O , make slightly alk. with N NaOH to litmus and sterilize. The soln. can be kept a long time but the reaction must always be alk. The film must be entirely dry. J. C. JUKRIENS

A simplified method for the determination of blood cholesterol. EMILY M. DAY AND ADOLPH BOLLIGER. *Australian J. Exptl. Biol. Med. Sci.* 7, 41-4 (1930).—The method requires less time and app. than others and gives a correlation of $\pm 5\%$ with that of Myers and Wardell, with 0.2 to 1.0 cc. of blood. Blood or plasma is dried at room temp. on two 7 cm. filter papers, folded, placed in 8 in. test tube and covered with CHCl_3 . The tube is then gently boiled for 15 min. in a beaker of CCl_4 and the CHCl_3 decanted and made up to 15 cc. with washings, from which 5 cc. aliquots are taken for analysis as in Myers' method, with 2 cc. As_2O_3 and 0.1 cc. conc'd H_2SO_4 . With 0.2 cc. of blood the drying may be omitted. C. G. KING

Rapid determination of blood serum protein. FERNAND KAYSER *Bull. soc. chim. biol.* 12, 533-5(1930).—To a mixt. of 25 cc 95% alc and 25 cc. acetone, add slowly with stirring 5 cc blood serum, let stand a few min., filter through a tared filter, wash the flask with 95% alc. and weigh. To sep. albumin from globulin, stir 10 cc. serum into a soln. of 80 g MgSO₄ in 65 cc distd. water, let stand 1 hr and make up to 125 cc. Filter and to an aliquot of the filtrate add 1 cc formal and 15 drops of 10% AcOH. Bring to boiling and filter, washing the ppt. with small portions of dil. AcOH until free from sulfate. Dry and weigh the ppt. and report globulin by difference.

C. G. KING

The determination of lactic acid in blood. J. LOISELUR AND R. MOREL. *Bull. soc. chim. biol.* 12, 538-40(1930).—The method is based upon the procedures of Sakowsky and Clausen (*C. A.* 16, 2342). To 3 cc fluoride plasma add 21 cc. distd. water and 2 g CaO, let stand a few min., add 6 cc satd. CuSO₄, let stand 1 hr with occasional shaking, and centrifuge. Det. lactic acid in the supernatant liquid by the Clausen method. Control expts. show that lactic acid can be recovered, that glucose does not interfere, and that the results compare closely with Clausen's when applied to different types of blood.

C. G. KING

Utilization of the Soja hispida seed freed from uricase in the detection and determination of allantoin. R. FOSSE, A. BRUNEL, P. DE GRAEVE, P. E. THOMAS AND J. SARAZIN. *Compt. rend.* 191, 1355-90(1930).—Under exptl. conditions the uricase of the soy bean was destroyed by heat and also by KCN without affecting the active properties of the allantoinase and uricase. The uricase-free substance was then studied for possible utilization in the detn. of allantoin in blood serum and in dog serum. The presence of allantoin in a diln. of $\frac{1}{100,000}$ can thus be unmistakably detected.

B. S. LEVINE

The use of uroselectan in urography. R. S. E. MORRAY. *U. S. Vet. Bur. Med. Bull.* 7, 111-6(1931).—A brief review of the developments leading to the evolution of Na 5-iodo-2-keto-1-pyridineacetate. Indications for its use and methods of administration and 3 case reports are given.

B. S. LEVINE

The alkaline decomposition of serine. FLOYD S. DAFT AND ROBERT D. COGILL. *J. Biol. Chem.* 90, 341-50(1931).—Serine is decomposed when heated in a strongly alk. soln., among the products formed being NH₃, glycine, alanine, (COOH)₂, and lactic acid. An intermediate decompn. product is pyruvic acid. Serine must be absent from solns. in which arginine is being estd. by alk. decompn., but this condition is secured by the pptn. of arginine with phosphotungstic acid. The methods commonly employed for the removal of NH₃ from protein hydrolyzates do not cause any significant decompn. of serine.

A. P. LOTTEROP

A reduction in the amount of blood required for the Folin micro method for blood sugar. HAROLD J. JEOHNS AND VICTOR C. MEYERS. *J. Lab. Clin. Med.* 15, 982-4(1930).—A modification is described of the Folin micro method (*C. A.* 22, 2762) for the detn. of blood sugar. The quantity of blood required is reduced to 0.025 cc., which may be measured by means of a capillary pipet.

E. R. MAIN

A routine blood chemistry unit. E. G. SCHMIDT. *J. Lab. Clin. Med.* 15, 1013-6(1930).—A mech. unit is described for the routine detn. of blood chemistries in the hospital lab.

E. R. MAIN

The preparation of some brominated oils and brominated esters. HELEN L. WIKOFF. *J. Lab. Clin. Med.* 16, 36-8(1930).—Methods are described for the prepn. of brominated oils and brominated esters suitable for use in röntgenologic work. Sesame oil contg. 40% Br appears to be the most suitable of the brominated oils. The brominated Me esters of the acids prepd. from cottonseed oil, which contain 42% Br, appear to be the most suitable of the esters. The latter have a low viscosity, are pale amber in color and more limpid than the corresponding oils or Et esters.

E. R. M.

Determination of copper in biologic material. J. M. INOUE AND F. B. FLINN. *J. Lab. Clin. Med.* 16, 49-51(1930).—An electrolytic method is described for the detn. of Cu in biologic material.

E. R. MAIN

Simultaneous determination of inorganic phosphate, sugar and lactic acid in blood. SERGIUS MORGULIS AND SHERMAN PINTO. *J. Lab. Clin. Med.* 16, 60-4(1930).—Procedures are described for the detn. of inorg. phosphate, sugar and lactic acid in the trichloroacetic acid filtrate obtained from 3 cc. of blood. The inorg. phosphate is detd. by the method of Kuttner and Cohen (*C. A.* 22, 606); the sugar, by a modification of the Hagedorn-Jensen method and the lactic acid, by the method of Friedemann (*C. A.* 23, 3245).

E. R. MAIN

Phosphorus metabolism. I. A system of blood phosphorus analysis. GUY E. YOUNGBURG AND MARIE V. YOUNGBURG. *J. Lab. Clin. Med.* 16, 153-66(1930).—

Detailed directions are given for the use of a colorimetric method for the detn. of P, in which the phosphomolybdate formed is reduced with SnCl_2 . The method is adapted to the detn. of total P, inorg. phosphate P, total acid sol. P and lipide P in whole blood, corpuscles and plasma. E. R. MARY

The determination of nonprotein nitrogen in 0.1 cc. of blood. ARTHUR K. ANDERSON and STACEY F. HOWELL. *J. Lab. Clin. Med.* 16, 183-6(1930)—A modification is described of the Folin-Wu method for the detn. of nonprotein N which requires but 0.1 cc. of blood. The pptg. agent is a mixt. of Na_2WO_4 and H_2SO_4 . The digestion is carried out in tubes graduated at 5 cc. E. R. MARY

A proposed chemical test for pyrogen in distilled water for intravenous injections. EDGAR B. CARTER. *J. Lab. Clin. Med.* 16, 289-90(1930)—One hundred cc. of the distd. water is heated to boiling in a clean pyrex beaker and treated with 10 cc. of 10% H_2SO_4 and 0.1 cc. of 0.05 N KMnO_4 . The color of the soln. should not be destroyed by boiling for 10 min. if pyrogens (fever producing org. substances of bacterial origin) are absent. E. R. MARY

Acetone as a yardstick for ketosis. ARTHUR T. BRICK. *J. Lab. Clin. Med.* 16, 291-3(1930)—A method is described by which the Legal test for acetone may be used as a quant. method for the estn. of acetone in the urine. E. R. MARY

Application of the quinhydrone method for the determination of the pH of solid medium. WM. STEENKEN, JR. *J. Lab. Clin. Med.* 16, 316-7(1930)—The quinhydrone method may be used for the detn. of the pH of agar or other solid media. Small portions (5 X 5 mm) are ground with 0.2 g. of powd. quinhydrone and the pH is detd. by use of the Cullen and Bülmann electrode (*C. A.* 10, 3101). E. R. MARY

A simple method of estimating "osmic acid," with some applications to cytological technique. R. PALMER. *J. Roy. Microscopical Soc.* 50, 221-6(1930)—Analysis of osmic acid is desirable, as indicative of deterioration and also of the amt. taken up by the tissue under treatment. Qual. tests for OsO_4 are available, of which Chapman's is best. A soln. of OsO_4 or a chlorosmate is heated with thionine in excess, acidified with HCl . A red color is observable in concns. greater than about 1:1,000,000. The reaction is applied to colorimetric estn., by prepng. a series of standards, which are in sealed tubes. 0.2 cc. of 2% OsO_4 , 32 cc. H_2O , 4 cc. of 5% thionine and 4 cc. of HCl (1% concn.) give 40 cc. of a soln. contg. 0.01% OsO_4 , from which other standards may be prep'd by dilution. For estn. of an unknown, 0.05 cc. is measured by a capillary pipet, mixed with 8 cc. H_2O , 1 cc. 5% thionine and 1 cc. of HCl (1% concn.), heated to boiling and cooled. The reaction is unaffected by HgCl_2 but gives a black ppt. if chromate-fixing agents have been used in the tissue. Data are given for the absorption of OsO_4 during the impregnation of frog kidney. Solns. of OsO_4 keep best in tightly stoppered bottles, so evapn. loss is minimized. C. W. MASON

Hyrax, a new mounting medium for diatoms. G. D. HANNA. *J. Roy. Microscopical Soc.* 50, 424-6(1931)—Resinous mixts. of anilope, S and HCHO may be prep'd, but these are rather deeply colored and do not harden readily. "Hyrax" is "a derivative of naphthalene" (no formula or recipe given) of pale straw color and resinous character $n = 1.82$, sol. in benzene or xylene, but not in H_2O or EtOH . It hardens by evapn. of solvent, it may be heated gently and is rendered fluid thereby, without discoloration. Prep'd slides and stained specimens have been permanent for several years, with only slight darkening if exposed to sunlight. Hyrax is very transparent to blue and violet. C. W. MASON

A method for the histochemical detection of iodine. U. HINTELMANN. *Z. wiss. Mikroskop.* 46, 486-7(1930)—The tissue is fixed in dil. HCHO and treated with 1% TiOAc . Yellow TiI is formed. Sections are best prep'd by the freezing method, and care should be taken not to dissolve out the TiI by subsequent staining treatments. The yellow color is seen best by reflected light against a dark background. TiOAc is better than Pb salts because no ppt. of chlorides or carbonates interferes, and is much cheaper than Pd salts. C. W. MASON

The fate of fructose in the animal organism. I. Determination of fructose by the diphenylamine method. W. W. ORRILL. *Biochem. Z.* 229, 85-99(1930)—From a detailed study of the method the following procedure for the detn. of fructose has been worked out. Place a mixt. of 1 cc. of the fructose soln., 0.1 cc. 20% alc. soln. of diphenylamine, and 1 cc. 25% HCl in a test tube for 20 min. in a vigorously boiling water bath. Cool the tube quickly in running water, ext. the color with 2-3 cc. isoamyl alc. and after sepn. remove the isoamyl alc. layer and dil. with 30 cc. alc., compare the color with that of a standard fructose soln. treated similarly. Quantities between 0.438 and 0.095 mg. can thus be detd. S. MORGULIS

Determination of the acid elimination in the urine. FRITZ MAINZER and MARCA

BRUHN *Biochem. Z.* 229, 216-32(1930).—Urine was collected under paraffin oil and analyzed immediately. The CO_2 content was detd. in the Haldane app. after 15 cc. urine was equilibrated in a tonometer at 38° with a gas mixt. of known CO_2 content. The total CO_2 was detd. in the Van Slyke app.; the H ion concn. was detd. with the quinone electrode, while the titratable acidity was detd. by titrating under oil to pH 7.45 with phenol red as an indicator. Not only the H ion concn. but the titratable acidity as well depends upon the CO_2 tension of the urine. At acid pH the error in detg. the urinary acidity is small, but at $\text{pH} > 8.0$ it may become 90-100% of the titration. With the aid of CO_2 tension and total CO_2 values satisfactory corrections may be obtained. On the other hand, the total acid excretion (acid + NH_3 - HCO_3) is largely independent of the abs. value of the CO_2 tension, provided the total CO_2 and acid are detd. under the same CO_2 tension.

S. MORGULIS

Analytical studies. XII LUDWIG PINCUSSEN *Biochem. Z.* 229, 233-7(1930), cf. C A 24, 4801.—A small extn. app. is described for extg. fat from liquids. A method is discussed for detg. Br in urine. Evap. to dryness 5-10 cc. of urine treated with 6-12 drops NaOH , then incinerate in a Ni crucible over an open flame. Filter the ash soln. into a special distn. app., acidify strongly with H_2SO_4 and mix with 20 cc. 10% KHSO_4 . Add a 2% soln. KMnO_4 , drop by drop, from a funnel until a permanent pink color remains and suck cold air through the app. for 1.5 hrs., passing it through 2 wash bottles with about 20 cc. 2% KI soln. At the close of the aeration titrate the KI solns. with 0.01 N $\text{Na}_2\text{S}_2\text{O}_4$, 1 cc. of which = 0.799 mg. Br. Another method described deals with the microdetn. of NH_3 and urea by the urtase procedure. The removal of NH_3 by aeration is facilitated by adding NaCl , which diminishes its soly. To 2 cc. urine add 1 g. NaCl , phenolphthalein and paraffin oil, then 2 cc. 33% Na_2CO_3 . Distil at 45° for 20 min. when theoretical results are obtained. In the detn. of urea the addn. of the NaCl does not have any significance.

S. MORGULIS

The manometric carbon dioxide determination according to Van Slyke. FAITZ MAINZER *Biochem. Z.* 229, 311-4; *Klin. Wochschr.* 9, 2401(1930).—A 50-cc. pipet is described so constructed that the gas vol. can be read at 0.5, 2.0, 4.0 and 6.0 cc. in the Van Slyke manometric CO_2 app. This is very useful in CO_2 detns. of urine, which often cannot be made in the ordinary app. calibrated only at 0.5 and 2.0 cc.

S. MORGULIS

Microdetermination of calcium and phosphorus in blood and tissues. G. WIMMARK AND B. VARLQVIST *Biochem. Z.* 230, 245-52(1931).—Blood, plasma or tissue fragments contg. 0.1-0.2 mg. Ca or 0.05-0.2 mg. P are digested with 0.5-0.75 cc. H_2SO_4 in a small Kjeldahl flask, and when the material is completely charred a few drops of concd. HNO_3 is added and the heating continued until a clear, colorless soln. results. The strongly acid soln. is transferred to a Pt dish, washed with H_2O and the soln. evapd., and then ignited to drive off all H_2SO_4 . The residue is taken up in H_2O , transferred to a beaker and warmed on a water bath. One cc. of satd. $(\text{NH}_4)_2\text{C}_2\text{O}_4$ and a drop of methyl red are added, and the reaction is adjusted to pH 5 with NH_4OH and AcOH . The pptn. is allowed to proceed at least 1 hr. and the material is filtered through a special porcelain rod with suction. The ppt. is washed twice with H_2O . The ppt. is dissolved in 0.5 cc. HCl , heated to 60° and titrated with 0.01 N KMnO_4 . The Ca can be detd. in 2 cc. plasma with an accuracy of 2-3% by this method. For the P detn. the digested material is diluted with 15 cc. H_2O and boiled for 3 min. To this are added 5 cc. 50% NH_4NO_3 and, after renewed boiling, 1 cc. 10% $(\text{NH}_4)_2\text{MoO}_4$. The mixt. is kept for 2 min. at 80° , and the ppt. is allowed to settle out for an hr. It is filtered again by means of the special porcelain rod, and washed 3 times with ice-cold water. An excess of 0.04 N NaOH is now added and titrated with 0.04 N H_2SO_4 against phenolphthalein.

S. MORGULIS

Microchemical demonstration of lead and mercury in the organism. H. BÜLL *Biochem. Z.* 230, 299-303(1931).—The Pb is detd. as $\text{K}_2\text{CuPb}(\text{NO}_3)_4$, which forms a beautiful cryst. ppt. The reaction is sensitive to 0.05 mg. Pb and can be carried out even in the presence of Cd or Hg in a concn. 300 times as great as that of the Pb. Blood, urine or feces is charred at a very low temp., and the charred mass is rubbed up with about 4 times its wt. of a mixt. of equal parts of KNaCO_3 and of KNO_3 . This is now fused in a porcelain crucible, dissolved in 2 N HNO_3 and filtered, if necessary. The soln. is evapd. and the nitrate is decompd. by evap. several times with 2 N HCl . Finally, the residue is taken up in 10 cc. 0.1 N HCl , warmed to 60° and H_2S is bubbled through it for 30 min. Then following neutralization with 10 cc. 0.1 N NaOH the H_2S is passed for 15 min. longer. After standing for 24 hrs. the ppt. is filtered off, washed with 1% NH_4NO_3 satd. with H_2S , and dissolved in hot dil. HNO_3 . After evapn. to dryness the residue is taken up in 3 drops of hot 0.1 N HCl , transferred to a microscope

slide, evapd over an open flame, and treated with a drop of 2% Cu acetate soln. Upon evapn the residue is treated with a drop of a mixt. consisting of equal amts of AcOH, satd NH_4OAc and H_2O to which an equal vol. of satd KNO_3 has been added. In the presence of Pb small brown or large black crystals appear. Pb is always found in feces where Pb had been ingested, and is never present in the blood unless it has been found in the urine or feces. Hg was demonstrated in the vomitus of poisoned persons by first heating this with concd HCl and KClO_4 until a yellow color developed, when a 3 mm piece of Cu wire is inserted for 24 hrs. The wire is washed, dried and sealed into a capillary tube. The Hg is distd off by heating the tube and the droplets of Hg are examd under the microscope. For urine, this must first be heated with concd. HCl and KClO_4 , satd with H_2S and the filtered off ppt. is dissolved in HCl and a crystal of KClO_4 . This is placed on a microscope slide, carefully evapd and treated with a drop-let of a reagent consisting of $\text{Co}(\text{OAc})_2$ and satd. NH_4SCN . Instead of this a small crystal of a mixt. can be added obtained by evapg an aq soln of 2 parts $\text{Co}(\text{OAc})_2$ and 5 parts of NH_4SCN . Blue crystals of mercuri-cobalt thiocyanate are formed in the presence of Hg. S MOACULIS

Microdetermination of the reduced and total glutathione in the liver. JOACHIM KÖHNHAU *Biochem Z.* 230, 353-72(1931).—Weigh quickly on a torsion balance about 0.5 g. of fresh liver (or liver kept on ice for not over 1 hr.), cut into several pieces and rub up with a little ignited quartz sand to a fine suspension with 9 cc. 10% $\text{CCl}_3\text{CO}_2\text{H}$. After leaving this for 5 min. filter the material through a double thickness of Schleicher and Schüll No. 595 in a Büchner funnel. Rinse the mortar with 4 cc. of the $\text{CCl}_3\text{CO}_2\text{H}$ and pour this also through the filter, repeating the washing once more. Pour the fluid into a 100 cc. Erlenmeyer flask. To this add 1 cc. 25% KI and 2 cc. 9.005 N I_2 soln. and after 2 min. titrate with 0.002 N $\text{Na}_2\text{S}_2\text{O}_3$, using 2 drops of a 1% starch soln. as indicator, make a blank titration, using 10 cc. $\text{CCl}_3\text{CO}_2\text{H}$, the difference multiplied by 0.02 giving the GSH (reduced glutathione) content in mg. With this procedure the error is $\pm 4\%$. S MORCULIS

The degree of oxidation within the animal body under the influence of different conditions. HIROSHI KIMURA *J. Biochem (Japan)* 12, 351-60(1930).—Müller's method (C A 24, 393) for detg the "Vakat O_2 " has been used with certain modifications. This method gives a measure of the amt. of O_2 needed to oxidize completely the org. substances excreted through the urine. It consists in oxidizing the urine with H_2SO_4 and KIO_4 at a high temp. and detg. in an aliquot the I_2 set free with KI. In another aliquot the N is detd. by the usual procedure. Also a correction for Cl must be made. 1 mg. Cl is equiv. to 1.207 mg. KIO_4 . K. points out that it is necessary first to heat the contents of the flask for 2 hrs. at 180° . The use of 0.5 g. KIO_4 for each 10 cc. urine is recommended, but of herbivorous urine only 5 cc. is taken. Human urine should best be dild. 2-3 times with water. Removal of the thyroid gland causes a diminution in the O_2 consumption, but the degree of oxidation is not altered. Feeding thyroid gland increases the amt. of oxidation, while the degree of oxidation is decreased far below normal. Adrenalectomy diminishes both the N and the total metabolism, but the degree of oxidation is not materially affected. S MOACULIS

A new microdetermination of bile acids in the bile. SATOSU NAKAGAWA AND HOSAHARU FUJIKAWA *J. Biochem (Japan)* 12, 399-410(1930).—When 0.5 cc. of an aq. soln. of bile salts is heated for exactly 2 min. in a vigorously boiling water bath with 2 cc. concd. H_2PO_4 and 2 drops of a 0.6% furfural soln. in abs. alc. a beautiful red color develops suitable for colorimetric study. Prep. a standard from 2 solns., one of which contains 0.01 g. scarlet and 1 g. HgCl_2 in 100 cc. and the other 1 g. $\text{K}_2\text{Cr}_2\text{O}_7$ in 100 cc. Mix 5 cc. of the first and 15 cc. of the second soln. and dil. with water to 30 cc.; this results in a soln. with a very permanent red color. The color is proportional to the concn. of the bile salts within the range of 5-20 mg. %. This red coloration is probably due to cholic acid, the other bile acids reacting very feebly or not at all. Fats, lecithin, cholesterol and alc-ether exts. of the brain give neg. results. S MOACULIS

Methods for the determination of total fixed base, sodium and potassium in urine. A FÖLLING *Skand Arch Physiol* 61, 27-31(1931).—For the total base detn. place 2 cc. of urine in a silica dish 7 cm. in diam. and, after the addn. of 4-5 drops 50% H_2SO_4 , evap. on a water bath and ash at 600° . After cooling, mix this with 1 g. of wet stannic acid and 3 cc. 35% HNO_3 (make the stannic acid by treating 20 g. Sn shavings with 300 cc. 35% HNO_3 and washing the residue several times with water by decantation, keep the metastannic acid wet in a glass stoppered jar), stir the mixt. and evap. to dryness. When cool, treat it with 10 cc. of a soln. of 2 cc. HNO_3 in 100 cc. water, to 10 cc. of which add 1 cc. EtOH just before using. Stir the mixt. well and filter through ashless paper. Evap. 5 cc. of the filtrate with a drop of 50% H_2SO_4 in a Pt dish, heating

continuously at first, then all over to a bright red heat for about 10 min. Dissolve the ash in 15 cc. neutralized water (add 0.02 N NaOH to distill water until a persistent red color to phenol red results). If not all the H_2SO_4 was driven off the color turns yellow and can be brought back to red by the addition of 0.02 N NaOH which is then subtracted from the results. Det. the sulfate by the benzidine micromethod as usual. Det. Na and K separately by removing the Ca and Mg as phosphates through the addition of NH_4OH . To 10 cc. urine add 10 cc. dil. NH_4OH and, after 1 hr., filter the mixt. Evap. 1 cc. of the filtrate in a silica dish to drive off NH_4OH , then add 5 drops of 50% H_2SO_4 and carry out the detn. as before for the total base. To another 10 cc. filtrate obtained from urine treated with dil. NH_4OH to remove Ca and Mg add 10 cc. 1% $CeCl_3$. Mix well, filter after 15 min. and treat a 4 cc. aliquot to det. the total base. Det. the Na alone as the complex orange Zn Na acetate salt. To 5 cc. urine add 5 cc. of a $UO_2(OAc)_2$ soln. contg. 27 g. of this salt and 3 cc. glacial AcOH in 1 l. of water. Filter off the phosphates after 16 min. To 2 cc. of the filtrate in an 18×150 mm. test tube add 4 cc. of reagent made by dissolving 20 g. $UO_2(OAc)_2$, 40 g. crystal $ZnSO_4$ and 7 cc. glacial AcOH in 180 cc. water and filtering after 2 days. This soln. should not give a ppt. when mixed with an equal vol. of alc. To the soln. in the test tube add 5 cc. 90% alc. with const. stirring and, after 20 min., filter through a Jena Gierste glass filter. Wash the tube and the filter 4 times with 1 cc. alc. Now dissolve the ppt. and wash back quantitatively into the test tube with hot water. Transfer to a small Erlenmeyer flask and titrate to a definite red with 0.1 N NaOH, using 5 drops 1% phenolphthalein. The cc. 0.1 N NaOH/8.77 = 0.1 milliequivalents Na in 1 cc. urine. Det. the K by difference from the K + Na detn. S. MORRIS.

Determination of ammonia in urine by extraction. GRITA HAMMARSTEN. *Scand. Arch. Physiol.*, 61, 48 (3(1931)).—The Whlmark rocking extractor is used. The NH_3 is liberated from 15–25 cc. urine by 5–10 cc. N NaOH and the liberated NH_3 is absorbed in a measured amt. of 0.1 N H_2SO_4 . The solns. in the 2 communicating chambers are covered with layers of acid free ether. Half of the NH_3 is absorbed in 30 min. so the shaking is carried out for that period of time and the standard acid is titrated with bromocresol purple as indicator. S. MORRIS.

Colorimetric determination of glucuronic acid in urine of man and rabbits. H. A. OOTA AND T. YAMANOUCHI. *J. Pharm. Soc. Japan* 50, 1039–75 (1930), German abstr. 135–8.—The method of detg. glucuronic acid in aq. soln. (cf. C. J. 23, 1718) was applied to the detn. of glucuronic acid in urine of man and rabbits. P. I. N.

The association of biochemical constituents and of certain cellulose esters. JEAN FRISLEUR AND LION VELLUZ. *Compt. rend.* 102, 151–61 (1911).—A method is described for the prepn. of a clear, solid membrane made from proteins, glucides or lipins. It is accomplished by using a carefully selected ratio of certain solvents. 1 g. in the prepn. of a casein membrane, 0 g. of dry acetate of cellulose is dissolved in a mixt. of 10 parts (17.0 g.) of acetone and 7 parts (33.1 g.) of abs. MeOH. 1 g. casein is dissolved in 0 g. of formic acid. The 2 solns. are mixed with rapid stirring, then poured on a plate till the solvents evap., when a clear membrane is obtained. By this method, membranes may be made from gelatin, casein, glycogen, gums, glycerides and other biochem. constituents. N. M. NAYLOR.

Biological reactions applied to the medico-legal identification of muscles. C. SIMONIN. *Ann. m'éd. légale criminol. police* 11, 32–7 (1911).—Pitg. serums can be used for the identification of fresh muscle only if they possess a very high activity, and only pos. reactions are conclusive. Expts. with 84 guinea pigs are described and showed that: (1) raw muscle behaves in the same way as blood in regard to anaphylactic reactions, (2) on holding the muscle loses part of its sensitizing properties. A. P. C.

The quantitative assay for the testicular hormone by the comb. growth reaction. T. F. GALLAGHER AND FRED C. KOER. *J. Pharmacol.* 40, 337–49 (1930).—Individual variations in response of Brown Leghorn capons to the injection of testis hormone were observed; they were not due to variations in age or wt. of the birds or original size of the comb. The best quant. results were obtained by using a min. daily dose required to produce an increase in growth of 3–7 mm. in 5 days. It is suggested that the strength of the hormone prepn. be expressed in bird units; a unit is defined as the amt. of hormone which, when injected in daily doses for 5 days, yields an av. increase of 5 mm. in length and height of the comb in at least 5 birds. C. R. MUR.

The specificity of the phenol reagent for the selective determination of tyrosine in proteins. VIRGINIA CROCATUS. *Compt. rend. soc. biol.* 101, 108–9 (1929), cf. C. J. 21, 3210.—The 1 dm. Jones phenol reagent provides a method for the estm. of tryptophan and tyrosine. An objection to this method for tyrosine has been that the reaction is not sp. In approaching this question from the point of view of the speed

of the reaction, C first studied the possible interference of the H_2SO_4 . Conclusions: The mercuric tyrosine compd (1) causes the limits of proportionality to vary greatly, and the H_2SO_4 diminishes the intensity of the reaction. If the solids which contain (1) are boiled for $\frac{1}{2}$ hr with HCl , the proportionality becomes perfect again, but the intensity of the coloration remains feeble. The use of H_2S gives the same excellent results as with tryptophan, but for a complete sepn. of tyrosine the flask satd. with H_2S should stand stoppered for at least 12 hrs before filtering; 5 cc of filtrate is boiled in a 100-cc. graduated flask for 30 min on the water bath, cooled and treated with the reagent. One mg tyrosine, without H_2SO_4 , serves as the reference standard. The proportionality is perfect. It is thus possible to det 0.5-4 mg tyrosine. J. D. S.

A new clinical method for the estimation of protein in urine. PHYLLIS M. TOORRY KERRIDGE. *Lancet* 1931, 1, 21-2.—The urine is mixed with a C suspension and C and protein are simultaneously pptd. with trichloroacetic acid. The grayness of the ppt. depends on the amt. of protein, when the amt. of C is const. The tone is compared with standards. The error is 0.025-0.05 g protein N%. F. B. SETBENT.

Triketohydrindene hydrate as a reagent for albumin, peptone and ammo acids. HERMANN GARDNER. *Lancet* 1930, 11, 525-6.—The color reaction with anhydridin is due to the presence of traces of free NH_3 or of NH_4^+ ions and is masked by NH_4OH or alkali in excess, and is, therefore, not a reliable reagent for detecting the products of hydrolysis of protein. Water contg 0.00865% NH_4^+ or in 1 cc. 0.77 mg NH_4^+ N, gives a blue coloration. F. B. SETBENT.

Optimum use of material in biological assays. WARREN W. NICHOLAS. *Bur Standards J. Research* 6, 77-87 (1931).—If a "characteristic curve" (cf. Trevan, C. A 21, 2913) representing the effect of increasing amts. of an agent on an organism is known approx., it may be used to det. the actual points on the curve accurately and efficiently. Such a curve is obtained by plotting the quantity or concn. of an agent, such as x-rays, ultra-violet light, toxins, inorg. drugs, hormones, vitamins, specific sera, etc., against the percentage of cases in which a given effect such as mortality is produced. The practice of using the same no. of animals to det. each point on the curve is inefficient, as the various points are then detd. with different degrees of accuracy. The point which will be detd. with least accuracy is that at which 0.5 the no. of animals show the given effect. To det. all points with the same degree of accuracy it is necessary to use a max. no. of animals for this point, and a decreasing no. for the points departing from this in either direction. A min. no. of animals will be used where all or none produce the given effect. To det. any 1 point independently, an expt. may be conducted with a small group of animals. From the result it is possible to calc. the value for the given point with the aid of an approx. curve, as well as the probable error of this value. The expt. may be repeated until the probable error lies within the desired limits. CORNELIA T. SWELL.

The study of undetermined carbon and glucidic carbon in normal urines. PAUL FLETCHY AND PIERRE AMBERT. *J. pharm. chim.* [8], 12, 350-64 (1930).—The main object was to find a relation between the quantity of glucidic C and undetd. C. Glucidic substances are considered those which not only are pptd. by the Cu-Ba reagent (C A 24, 1625) but also are reducers of Hg salts (F. and Marqne, C. A 24, 591, 559). From a sample of urine 3 liquids were prepd. at near 0° for examn.: (1) the urine defecated with Hg salts, (2) the Cu-Ba pptn. filtrate of (1) and (3) the H_2SO_4 soln. of the ppt. In each of the 3 liquids were detd. (a) total C (Nicoloux, C. A 23, 628), (b) total N, (c) urea, (d) Hg reduction, (e) Cu reduction (F. and Boutot, C. A 16, 4222) before and after hydrolysis. Conclusions: C and N contained in the defecate were completely recovered in solns. (2) and (3), while Hg reduction showed a loss of 40%, Cu reduction 20.9% before and 25.3% after hydrolysis. Therefore the losses of reducing power must be due not to removal by pptn., but to instability of the reducing substance. As the Cu-Ba ppt. earned down only 20% C in the form of truly glucidic C (0.710 g per l. = 1.78 g dextrose per l.) while the defecate by Cu reduction after hydrolysis showed only 0.717 g dextrose, i. e., about half as much, the heretofore undetd. quantity of "nonreducing glucide" must be of about the same order as the reducing quantity. In addnl. expts., starting from the Cu-Ba ppts., the liquids showed no appreciable rotatory power. They contained Cu-reducing substances, exclusively aldoses (no ketoses), and with PhNH_2NH_2 gave crystals seemingly identical with the glucosazones of Grimbart and Bernier (C A 4, 1497); however, they differ greatly in m. p. S. WALDBOTT.

Determination of biliary salts in bile. I. Gasometric determination in one process. II. Volumetric determination by formal titration. III. Evaluation of biliary sulfur. L. COTTEY. *J. pharm. chim.* [8], 12, 485-98 (1930); cf. C A 21, 2914.

3211, 22, 2580-4550—1. To avoid preliminary gasometric detn of N in the free amino acids, urea and NH_4OH present, remove these by NaBrO , which liberates their N without affecting the mols of glycine and taurine cholates. Defecate the bile with hot 95% alc, treat an aliquot of the filtrate with NaBrO , remove excess of this with H_2O_2 , hydrolyze the bile salts with NaOH , neutralize with HNO_3 and det the vol of N as usual in a ureometer by means of NaNO_2 and dil HNO_3 . Calc. the results in terms of g of N per l of bile. Make a blank test on the reagents used without bile, deducting any gas vol formed. II. Treat the bile as before with 95% alc., NaBrO and NaOH , then eliminate cholalic acid and carbonates by dil HCl , filter, neutralize with NaOH (phenolphthalein as indicator), add BaCl_2 , filter again and acidify slightly. Compare a definite aliquot corresponding to 3 cc of bile (in tube B) with an equal vol of a blank detn (tube T) prep'd exactly as B except that bile is omitted. To both B and T add 6 drops of phenolphthalein and $\text{Ba}(\text{OH})_2$ soln of exactly known strength ($C \times 0.1 N$) to pale pink and 10 cc of neutralized CH_3CO . Now add to T 0.5 cc of $\text{Ba}(\text{OH})_2$ soln to bring the soln to dark red, then det in B the vol (n cc.) of $\text{Ba}(\text{OH})_2$ soln necessary to produce the same tint as in T. The content of N (g per l of bile) is then $N = (n - 0.5) \times C \times 0.0014 \times 1000/3$. III. An adaptation of the benzidine sulfate method (C. A. 21, 2145, 22, 4140) is used. After defecation with alc. and evapn., destroy the org. matter by treatment with HNO_3 and then H_2O_2 . Hydrolyze the filtered-off benzidine sulfate (obtained by pptn in an acetone medium) with boiling water and while hot titrate the free H_2SO_4 with 0.02 N NaOH in a bromothymol blue soln as indicator until a permanently green tint is formed, to which the indicator soln was previously adjusted. Each cc. of 0.02 N NaOH indicates 0.00032 g S. C. prefers the above method I although the capacity of the instrument limits the size of the sample used for analysis. Method II requires experience to read the end points, and double filtration is a disadvantage. Finally, C. quotes the results of Escalardi (*Polskiego Archiwum Medycyny i Farmacji* 6, 4(1925)) on comparative detns. of biliary acids in org. liquids by colorimetry (Herrfeld and Haemmerli, C. A. 19, 2965), fluorescence (Raue, C. A. 21, 2710) and gasometry (Churray and C. C. A. 21, 3211).

S. WALDROTT

Determination of cholesterol in blood. E. CANALS. *Bull. pharm. Sud-Est; Schweiz. Apoth.-Ztg.* 68, 49-50(1930).—The av. normal content of blood serum and plasma in cholesterol is 1.5 g per l. The colorimetric method of detn by Grigaut (cf. C. A. 4, 2513, 2514, 6, 497, 1447, 8, 147) is only approx. accurate. The wt. method by Grigaut for total cholesterol, or better still the technique of Windaus for free and combined cholesterol, is recommended. Details are given for all methods. S. W.

Clinical determination of the albumin-globulin ratio in spinal fluid. WU G. EXTON AND ANTON R. ROSE. *J. Am. Med. Assoc.* 96, 36-7(1931).—For the detn of total protein equal vols. (1 cc.) of spinal fluid and Exton's protein reagent are mixed and the resulting turbidity is measured in the scopometer. For the detn. of globulin, equal vols. (2 cc.) of spinal fluid and a cold soln. of $(\text{NH}_4)_2\text{SO}_4$ are mixed and let stand a few min. for the globulin to salt out. The mixt. is filtered and the filtrate passed through the same filter until clear, care being taken to drain the test tube thoroughly. The filtrate is an albumin soln. and may be used later if it is desired to ppt. and measure the albumin separately. The salted out globulin on the filter paper is placed in a clean test tube, 2 cc. dist'd H_2O poured on the filter paper to bring the globulin into soln. and this is passed thrice through the same filter paper to insure a perfect homogeneous globulin soln. After the funnel has been thoroughly drained, 0.2 cc. dist'd H_2O is added in the tip of the filter cone to wash down the globulin that adheres to the tip of the cone. One cc. $(\text{NH}_4)_2\text{SO}_4$ soln. is added to the globulin soln. and mixed. The globulin is pptd. by adding 2 cc. of a cold aq. soln. of phenol and mixed by gently inverting the tube once. The globulin phenol ppt. is allowed to stand at least 15 min. and then the turbidity is read in the scopometer sometime between 15 and 30 min. after pptn. Mg globulin per 100 cc. of spinal fluid is obtained by dividing the junior scopometer scale reading by 4. Albumin is detd. by subtracting the globulin from the total protein. If a sep. pptn. of albumin is desired, 1 cc. dist'd H_2O is mixed with 2 cc. of the original $(\text{NH}_4)_2\text{SO}_4$ filtrate and pptd. with phenol as for globulin. The no. of mg. of albumin per 100 cc. of spinal fluid is then obtained by dividing the junior scopometer scale reading by 2.

R. C. WILLSON

TAILLANDIER, M.: Spectrophotométrie et photométrie appliquées à l'analyse biologique. Paris: N. Maloine. 256 pp. F. 40

C—BACTERIOLOGY

CHARLES H. MORRIS

The effect of iodine on microbes. I JADIN. *Compt. rend. soc. biol.* 105, 481-2 (1910) —Suspensions of *E. typhi* (I), *paratyphosus B* (II) and *paratyphosus C* (Hirschfeld's bacillus) (III) were heated on the steam bath at 56° for 30 min. The suspensions were iodized at 50° for 2-3 hrs. by adding first several drops of Na_2CO_3 soln., then I iodide soln. until free I_2 was present. The excess I_2 was removed by centrifugation, and the iodized microbes were injected into rabbits to get immune sera. The antigen from the injection of iodized I agglutinated both normal and iodized I (1:4000) but not II or III. Similar specificity was shown by II and III. Iodized microbes, unlike serum proteins, preserve with their antigenic properties all their specificities. B. C. BRUNSTETER

The action of lipid solutions upon the growth of acid-fast bacilli. TEOPHIL NYRÉN. *Beitr. Klin. Tuberk.* 73, 238-50 (1929) —Lecithin from 5 different sources (human, cow, pig, plant and egg) was used in disperse soln.; it favored the growth of tubercle bacilli. Cholesterol had no stimulating influence. Human lecithin retarded the growth of saprophytes. There was no inhibition from lecithin upon cold blood tubercle bacilli and saprophytes. A no. of the acid fast bacilli grew upon the lecithin nutrient medium with marked granulations. H. J. CORPER

Biological chemical observations on the tubercle bacillus. V. A. VASSILIOV. *Z. Tuberk.* 54, 219-26 (1929) —Tubercle bacilli were extd. with chloroform and a wax-like dark fraction was obtained in addition to a second fat-like fraction. Various substances were tested to det. the soly. of these fats. The chem. investigation revealed that the m. p. could not be detd. because carbonization occurred at 250°. Extn. of the substances with ether revealed an absence of fat and the constituents of fat such as glycerol, elain, stearin, etc., were absent. Conclusion: The so-called fat and waxy substances of the tubercle bacillus are not related to fat or wax. The watery extn. of the chloroform emulsion disclosed an aldopentose, however. H. J. CORPER

The chemistry of disinfection. WILDER D. BANCROFT AND G. HOLMES RICHIE. *J. Phys. Chem.* 35, 511-30 (1931) —Disinfection is explained on the basis of the colloidal nature of cells. The first stage of decreasing stability of the cell colloid is associated with the phenomena of stimulation, the next stage of reversible coagulation inhibits the activity of the organism (antiseptics), and the final stage of irreversible coagulation brings death to the bacteria (disinfection). These colloidal changes of yeast cells and of certain bacteria were observed experimentally with the ultramicroscope. The chem. concept of disinfection is shown to be unsound because of the inapplicability of the mass action law, the lack of stoichiometrical relations and the abnormal temp. coeffs. Conclusion: Disinfection consists first, of an adsorption of the drug and, secondly, of the coagulation of the cell colloids. MALCOLM DOLE

The reaction of purines to protein in microorganisms. ÉMILE F. TERROINE AND I. SZUCS. *Ann. physiol. physicochim. biol.* 6, 123-49 (1930) —The content of total N cannot be taken as a measure of protein in microorganisms. It is, however, true that a given species of bacterium always contains the same purine/protein N ratio. H. EAGLE

Influence of starvation (complete starvation and nitrogen starvation) on the proportion of proteins and of purines in microorganisms. ÉMILE F. TERROINE AND I. SZUCS. *Ann. physiol. physicochim. biol.* 6, 157-77 (1930) —Either complete or N starvation causes a decrease in the total protein of yeast and of *St. nigra*. Purine and protein N fall concomitantly at first, in later stages, the decrease in protein N is proportionately larger. H. EAGLE

The viability of *B. pestis*. I. S. TINKER AND G. P. REDNEW. *Arch. Schiff-Tropen Hyg.* 34, 554-9 (1930) —*B. pestis* is not affected by Hayem's soln. or by Grams stain 1:10. It is, however, killed by a modified Turk's soln., or by fixation of the smear with EtOH for 5 min.-2 hrs., followed by Grams stain. The reagent used for hemoglobin detn. kills it in 45 min. Dried and stained, the organism remains viable for 2 hrs. H. EAGLE

The vibriocidal powers of the water of certain rivers of India. SARANJAM KHAN. *Indian J. Med. Research* 18, 361-5 (1930) —Water of the Ganges and Jumna rivers has vibriocidal properties which are completely destroyed by heating at 55° for 1/2 hr. H. EAGLE

The bacteriological diagnosis of tuberculosis. J. MOURIZ. *Klin. Wochschr.* 9, 2249-51 (1930) —Hahn's method was found to be the best for isolating tubercle bacilli from mixed culture. No more than 4-5% H_2SO_4 should be used, a potato-egg medium

was found to be the best. A special tube was devised to keep the medium moist

H. EAGLE

Formation of hydrogen peroxide by lactic acid bacteria. ALFRED BEATHO AND HANS GLÖCK. *Naturwissenschaften* 19, 88(1931).— H_2O_2 is formed in oxidative metabolism by the facultative, anaerobic, catalase-free lactic acid producers *Bact. delbrücki*, *B. sugartii* and *B. acidophilus*. The O_2 consumption was followed in Barcroft manometers with glucose as "donor." H_2CN , 0.0005 mol., was without influence. The H_2O_2 was detected with Ti sulfate and sited as Ce peroxide (Wieland and Rosenfeld, *C. A.* 24, 2155) or by direct titration. Generally the theoretical amt. from $2H + O_2 = H_2O_2$ was found. Catalase addn. reduced the amt. of H_2O_2 to half. If the H_2O_2 was not bound during the reaction, its increasing concn. damaged the cells and resulted in decreased total production. The conditions were about as follows: 3 cc. bacterial suspension of 10 mg. dry substance, 3 cc. 0.1 M glucose, 3 cc. 0.2 M borate buffer of pH between 8 and 9, 1 or 2 cc. 0.002 M cerous sulfate were added together to give an initial pH of 8.2 to 8.3. For *Streptococcus casei*, which also respire unhampered by H_2CN , no H_2O_2 formation was found.

B. J. C. VAN ORSSEL

The use of small additions of acid for increasing the germicidal action of E. C. on bacterial spores. C. S. RAM AYYAR. *Agr. J. India* 25, 213-0(1930), cf. *C. A.* 24, 5328.—Spores of refractory bacteria, such as *B. subtilis*, *B. albolactis* and *B. mesentericus*, were practically completely destroyed after 30 min. contact with a soln. contg. 0.2% of 1% C. (a special hypochlorite soln. contg. 2% of available Cl) and 0.05% of a N soln. of citric acid. With a 0.2% soln. of E. C. alone the spores were not destroyed even after 60 min. contact, although the no. was reduced to a considerable extent.

K. D. J.

Remarks on the composition of peanut meal peptone and its use for the culture of pathogenic bacteria. ALBERT BEATHO, G. AMOUREUX AND DENIS PÉTIT. *Bull. soc. chim. biol.* 12, 1029-30(1930).—Analysis is given of a pepsin (I) and a pancreatic peptone (II), illustrating differences depending upon the method of prepn. I gave good toxin production and II did not, although both gave good growth with *B. tetanus*.

C. G. KING

Physical difference among bacteriophages. F. M. BURNET. *Australian J. Exptl. Biol. Med. Sci.* 7, 27-35(1930).—Bacteriophages from *Salmonella*-dysentery groups of bacilli fell within 2 groups with regard to the difference in heat stability and extension of filter paper.

C. G. KING

Balanced salt action as manifested in bacteriophage phenomena. F. M. BURNET AND MARGOT MCKIN. *Australian J. Exptl. Biol. Med. Sci.* 7, 183-98(1930); cf. *C. A.* 24, 5787.—Bacteriophage (I) did in a soln. of Na, K and NH₄ salts were inactivated by heating 1 hr. at 60°, but the presence of Ca, Ba or Sr salts partially or completely prevented such action. In the absence but not in the presence of bivalent cations or broth, Janus green, toluidine blue and methylene blue inactivated a dil. soln. of I. Different types of I show characteristic individual differences in sensitivity toward inactivating and protecting substances. The effects related to the Na:Ca balance are like those of living protoplasmic structure. Electrophoretic expts. indicate a decrease in effective size of I in the presence of excess Ca ion.

C. G. KING

The electrical behavior of bacteriophages. F. M. BURNET AND MARGOT MCKIN. *Australian J. Exptl. Biol. Med. Sci.* 7, 199-209(1930).—In electrophoretic expts. 20 types (*Salmonella*, dysentery and staphylococcus) of bacteriophage all showed a negative charge. The migration rate and adsorption of proteins (globin and gelatin) varied greatly. In one case cathodic concn. was obtained for a type which acted only on rough strains and was not sensitive to methylene blue, toluidine blue or Janus green.

C. G. KING

Oligodynamic dilutions. JULIUS MEYER. *Chem.-Ztg.* 55, 85-6(1931).—The term "oligodynamic" as used in biology and medicine is a misnomer. It was used originally to designate germicidal telekinesis of metals such as Ag or Cu plates placed in aquaria or flower vases. At present under this term is included such germicidal action as is manifested by extremely high dilns. of metallic salts, despite the fact that the phenomenon of telekinesis is entirely absent. The dilns. with which the biol. and medical investigators are experimenting often are extremely high from the point of view of the analytical and phys. chemist. For example, K. König (cf. *C. A.* 22, 4177) worked with dilns. of $AgNO_3$ and $Pb(NO_3)_2$ varying from 10^{-12} to 10^{-14} . For $Pb(NO_3)_2$ in a 10^{-12} diln. 55,000 l. of the soln. contains 1 mol. of the reagent and for $AgNO_3$ 28,000 l. contains 1 mol. of the reagent. Results obtained from work with such dilns. are scientifically valueless since such dilns. cannot be distinguished from water of the highest purity.

B. S. LEVINE

Experiments on bacteriophage adsorption by vulnerable bacteria. CLAUDE P.

BROWN AND FRANK GEMAR *U S Vet Bur. Med Bull* 7, 224-9(1930)—The expts. show that it is possible to adsorb bacteriophage obtained from sewage by using a vulnerable organism (*E. coli*) against which it is actively lytic. Temp control is necessary in order to bring about the adsorption of bacteriophage when the organism against which it is actively lytic is used for this purpose. A temp as low as 25° still permits lysis while 10° or lower prevents lysis. B S LEVINE

Disinfection. II. The manner of death of certain bacteria and yeasts when subjected to mild chemical and physical agents. GEORGES KNAYSI AND MORRIS GORDON *J Infectious Diseases* 47, 303-17(1930)—The order of death of *E. coli* and of *Schizosaccharomyces pombe* when produced either by chem. or by phys. agencies is detd by the distribution of resistance among the cells. That this distribution is not uniform among a culture of microorganisms is shown by the structure of the death curves. The similarity in results obtained with *F. coli* and *S. pombe* does not justify the theory advocating a difference in the order of death between small and large cells. III. The taking up of iodine by yeast cells. *Ibid* 318-21—The cells of *Saccharomyces cerevisiae* take up I from its aq. solns according to the adsorption isothermal. These results are in agreement with microscopic observations which show that the colloidal structures of the cells and especially the membrane, take up I to a great extent and give up much of it on diln. When $HgCl_2$ is used instead of I, the adsorbed salt may be pptd as HgS . The membrane, with its high adsorptive power, becomes black, while the rest of the protoplasm becomes gray only after more prolonged standing. This suggests that the mechanism of disinfection is closely related to that of dyeing and tanning. JULIAN H LEWIS

The use of acid fuchsin in Russell's triple sugar medium. G D CUMMINGS *J Infectious Diseases* 47, 359-66(1930)—Undecolorized acid fuchsin is recommended for use in the Russell medium because of its chem. stability, ease of prepn., ability to give clear-cut reactions and lack of inhibitory effect. The use of undecolorized acid fuchsin would eliminate the variety of methods for the prepn. of Andrade indicator now to be found in the literature. JULIAN H LEWIS

The growth and toxin production of *Corynebacterium diphtheriae* in synthetic media. MARY E MAVER *J Infectious Diseases* 47, 381-98(1930)—Diphtheria bacilli can be adapted to grow and produce potent toxin in a modified Braun and Hofmeier synthetic medium. The chief modification was of the N-bearing constituents. The cystine content was increased fourfold, asparagine or NH₂, succinate replaced the Na aspartate, and glycine was added. It appears that the medium must provide not only nutrients for rapid growth and pellicle formation, but these nutrients should also favor the appearance or maintenance of the toxigenic variants that occur in consequence of a dissociative process. JULIAN H LEWIS

Hexylresorcinol in experimental tetanus. GEORGE E COLEMAN *J Infectious Diseases* 47, 410-5(1930)—Hexylresorcinol S T 37, when mixed with tetanus toxin in 16.6% soln. is capable of neutralizing 10 and less than 100 minimal lethal doses. It is bactericidal for the vegetating forms of tetanus bacilli within 0.5 hr. *in vitro*, but does not kill resistant spores in 12 hrs. When injected into the devitalized or necrosed tissue of guinea pigs infected with tetanus spores, it is without action in modifying the course of the disease. JULIAN H LEWIS

Decomposition of urea by *Proteus*. ALEXANDER A KAY, WM M GRIBBS, ARTHUR W WALKER AND RUTH E JUNG *J Infectious Diseases* 47, 490-502(1930)—The destruction of urea by *Proteus* in various media was studied. Urea does not meet the N requirement of the organisms and the inference is warranted that the action on urea is only incidental to the metabolism of the bacteria and is not concerned with its vital needs. The enzyme responsible for urea destruction is of endocellular nature since $CHCl_3$ -killed, washed and dried organisms are active. JULIAN H LEWIS

The antagonistic substances formed during bacterial fermentation. B KLEIN *Z Immunitäts* 66, 530-42(1930)—In culture media contg. carbohydrates bacteria form antagonistic substances that behave like antibodies in that they produce agglutination and bacteriolysis of whole bacteria and give precipitin and complement fixation reactions with bacterial exts. These reactions are, however, not sp. The agglutination is due to acids formed during fermentation. The antagonistic substances are formed not only in lab. media but also in natural and com. processes involving fermentation such as the curdling of milk and the making of beer, wine and cheese. J. H. L.

The antiseptic properties of uroselectan. RUGGERO ASCOLI *Minerva med* 1930, 11, 415-7—Uroselectan, the Na salt of 5-iodo-2-keto-1-pyridineacetic acid, in concn. of 10-20% retards the growth of *E. coli* and *Staphylococcus* in urine. The inhibition is

more marked at a concn of 30%, but even a concn of 40% fails to kill the organisms. The substance, therefore, has bacteriostatic but no antiseptic power. P. M.

Peroxidase in relation to bacterial growth with special reference to the influenza bacillus. LUCILE R. ANDERSON. *J. Bact.* 20, 371-8 (1930).—*B. influenzae*, hemolytic and non hemolytic, and *B. hemoglobinophilus-canis* do not produce peroxidase. *B. parainfluenzae* produces it under both aerobic and anaerobic conditions. Not one of 13 anaerobes tested produced it according to the benzidine reaction. Of 44 aerobes tested all but 4, two of which were streptococci, gave a positive reaction. Banana reacts positively to the benzidine test. The "X factor" of Fildes seems to be associated with peroxidase. JOHN T. MYERS

Micrococcus niger, a new pigment-forming anaerobic coccus recovered from urine in a case of general arteriosclerosis. IVAN C. HALL. *J. Bact.* 20, 407-15 (1930). JOHN T. MYERS

The effect of the oxidation-reduction character of the medium on the growth of an aerobic form of bacteria. W. P. ALLYN AND I. L. BALDWIN. *J. Bact.* 20, 417-40 (1930).—This factor exerts an important influence on the growth and behavior toward atmospheric O of certain aerobic bacteria (*Rhizobium*). In agar shake cultures these organisms establish thin zones of growth, the level of which depends on the oxidation-reduction character of the substrate. This also acts in part the ability to initiate growth. By reducing with appropriate concns of cystine a medium which is slightly too oxidized for these organisms, it is possible greatly to increase the no. of sources which develop into colonies in Petri dish cultures. Hence this is important in making plate counts. The organisms are able to make slight adjustments if the media are too oxidized. The introduction of sand, finely divided filter paper, etc., into a liquid medium which was too oxidized to permit growth overcame the condition, probably because of the trapping of small vols. of medium where the organism could make the necessary adjustments, as suggested by a lag period interpreted as time required for this adjustment. The bacteriostatic effect of certain oxidizing and reducing substances seems due in part to the poisoning of the medium at an unfavorable potential. Their toxicity depends largely on concn. In relation to the oxidation-reduction potential of the system. Similar concns. were inhibitory under one set of conditions and stimulatory under another. J. T. M.

A marine denitrifying organism. BLOWEN LLOYD. *J. Bact.* 21, 89-96 (1931).—A denitrifying organism was isolated from sea water off the Scottish coast, and denitrification curves were constructed. JOHN T. MYERS

The effect of temperature upon the production of hydrogen sulfide by *Salmonella pullorum*. RALPH P. TITSLER. *J. Bact.* 21, 111-8 (1931).—An incubation temp. of 30° is optimum, and temps. above 37° inhibit H₂S formation by *S. pullorum*. Slight quant. differences were noted between aerogenic strains at 30°. The ability to produce H₂S parallels the ability to produce gas from carbohydrates. Difco Pb agar is satisfactory for demonstrating H₂S formation. JOHN T. MYERS

A modified Loeffler's blood serum medium useful in the routine health department examination for diphtheria and streptococcus infections. LEON S. MEDALIA, KARL R. BAILEY AND CATHERINE ATWOOD. *J. Bact.* 21, 118-37 (1931).—Loeffler's medium is modified by the addn. of 4.5% of N NaOH, yielding a final pH of 7.6, bromothymol blue is used as an indicator. JOHN T. MYERS

An instance of pigment formation by *B. pestis*. A. BESSONOVA AND M. LOCHOV. *Zentr. Bakt. Parasitenk.* 1 Abt., 119, 35-8 (1930). JOHN T. MYERS

A polysaccharide-building streptococcus. J. OERSEKOV. *Zentr. Bakt. Parasitenk.* 1 Abt., 119, 88-93 (1930).—A substance analogous to the polysaccharides of the pneumococcus was found. JOHN T. MYERS

Bacteria which destroy blood pigment. (Studies with the help of bacteria which form hydrogen peroxide on heated blood plates. K. BINGOLD. *Zentr. Bakt. Parasitenk.* 1 Abt., 119, 97-100 (1930).—*Pneumococci*, *Streptococcus mucosus*, *S. viridans* and to a lesser degree *S. pyogenes* can split oxyhemoglobin, methemoglobin or hematin to substances of simpler constituents. This seems to be due to H₂O₂ formation. This can occur only in the absence or the presence of minute amts. of catalase. JOHN T. MYERS

Changes in virulence and hemolytic power produced in hemolytic streptococci by growth in hematin agar. K. BINGOLD AND K. BACH. *Zentr. Bakt. Parasitenk.* 1 Abt., 119, 100-3 (1930).—Growth in hematin agar decreases virulence and hemolytic power. JOHN T. MYERS

The pigments of timothy grass bacilli. ERWIN CHARGAFF. *Zentr. Bakt. Parasitenk.* 1 Abt., 119, 121-3 (1930). JOHN T. MYERS

The phenomenon of hemolysis through the action of bacteriophage. ELISABETH BECHTEL. *Zentr. Bakt. Parasitenk.* 1 Abt., 119, 340-9 (1931). JOHN T. MYERS

Cytological observations on the dissociation process in *B. pertussis*. M. POKROVSKAYA. *Zentr. Bakt. Parasitenk.*, I Abt., 119, 353-61 (1931) — A nucleus of chem. structure analogous to that in higher forms is demonstrable in certain phases of bacterial development. JOHN T. MYERS

The disinfecting action of simple organic mercury compounds. GERTRUD MEISSNER. *Zentr. Bakt. Parasitenk.*, I Abt., 119, 375-9 (1931) — There are halogen Hg alkyl compds. which have a higher killing power for staphylococci and anthrax bacilli than does HgCl₂. The growth inhibiting power of Me, Et, Pr, Bu and Am Hg compds. is about 25-1000 times higher than that of HgCl₂. The action of the higher members of the series is weak or nil. Cl and Br salts are about equal in action, but I salts are weaker. The killing action is also better than that of HgCl₂. The hydroxides have a much weaker action than the halogen salts. Anthrax spores were more susceptible than staphylococci. JOHN T. MYERS

The influence of alkaloids on the metabolism of *Bac. pyocyaneus*. T. H. AMARO. *Zentr. Bakt. Parasitenk.*, I Abt., 119, 463-72 (1931) — Cocaine, nicotine, physostigmine, pilocarpine and strychnine inhibit methylene blue reduction in proportion to their concn. Atropine and morphine inhibit only at moderate concns. (0.1 to 1.0%), not above or below. JOHN T. MYERS

The influence of plant growth on the activity of root bacteria. GERTRUDA GALT. *Zentr. Bakt. Parasitenk.*, II Abt., 82, 44-69 (1930) — The greatest activity of the N-fixing root bacteria is concurrent with greatest plant growth. N₂ production in soil is also greater during plant activity. JOHN T. MYERS

The fermentation of pentoses by certain propionic acid bacteria. MARIAN FOOTE, E. B. FARD AND W. H. PETERSON. *Zentr. Bakt. Parasitenk.*, II Abt., 82, 379-89 (1930) — Is common with the majority of organisms, the propionic acid bacteria do not ferment the pentoses as readily as the hexoses. Out of 20 cultures tested, 13 fermented arabinose and only one fermented xylose. The fermentation of pentoses as compared with glucose is marked by a decrease in the percentage of volatile acids and an increase in non-volatile acids. The ratio of propionic acid to acetic acid is decreased. The non-volatile acids consist of lactic and succinic. The latter was identified by its m. p. and the Ba content of its salt. The succinic acid is probably derived from the carbohydrate rather than from the yeast water. JOHN T. MYERS

The influence of high dilutions of quinine on the bacterial flora in hay infusion. M. FENNER. *Zentr. Bakt. Parasitenk.*, II Abt., 83, 63-8 (1931) — Quinine at a concn. of about 10⁻¹⁰ M favors the growth of *B. proteus*. *Paramecium* is stimulated at a concn. of about 10⁻¹⁰ M. JOHN T. MYERS

The lactic mannitic fermentation of sucrose. BOLCATO VIRGILIO. *L'ind. zaccarif.*, 22, 427-32 (1929) — Partial anaerobiosis is a needed condition for the activity of mannitic enzymes. A smaller amt. of mannitol with aerobic fermentation is compensated for by more lactic and acetic acids. Fermented juices defecated with basic Pb acetate gave, with Bertrand's liquid after 1 day, polarization of 7.60 and 0.528% invert sugar, after 2 days, polarization of 3.70 and 4.000% invert sugar; after 3 days, polarization of 2.20 and 3.040% invert sugar, after 4 days, polarization of 2.70 and 1.630% invert sugar. A variation in the quantity of invert sugar influences the saccharimetric value, rendering it unsuitable for a control of the fermentation process. This is best conducted by considering the principal products: lactic acid (45% of the sucrose), AcOH (12-5%), mannitol (30-5%). AcOH is detd. by adding H₂SO₄ and distg. in a current of steam. Lactic acid is detd. by subtracting AcOH from total acid. For detg. total acid the fermented liquid is neutralized with CaCO₃, in 1 cc. of the clear liquid Ca is pptd. with NH₄ oxalate and the oxalic acid titrated with 0.1 N KMnO₄. From this the Ca and lactic acid are detd. Mannitol is best detd. by Gayon and Dubourg's method of crystn. from alc. at 85° and weighing. The estn. is preceded by an alc. fermentation for eliminating the remaining sugars. The yeast has only a slight action on the mannitol. If samples are taken regularly from the beginning of fermentation it is noted that, while the lactic acid is formed continuously with a regular speed, the mannitol is all formed within the first 4-5 days. In the same time nearly all the AcOH is formed. The same juices with the sucrose inverted by warm H₂SO₄, filtered and submitted to fermentation, in 2 days show a change in the polarization from neg. to pos., denoting a more rapid consumption of levulose than glucose. The destruction of levulose is very rapid, as is the production of mannitol and AcOH, and continues until all the mannitol is formed. Despite the consumption of glucose (shown by the increase in lactic acid) the pos. polarization continues to rise and reaches its max. with the total disappearance of levulose and arrest in the production of mannitol. After this the pos. polarization is gradually reduced by the transformation of glucose into lactic acid. This is caused by

the high neg. value of the sp. rotary power of levulose, $[\alpha]_D = -92^\circ$, in comparison with the pos. value of glucose, $[\alpha]_D = +52^\circ.5'$, and the lower fermentation speed of glucose with regard to levulose.

The indole titer of Gersbach. E. L. KRUGERS DAGNEAUX *Chem. Weekblad* 28, 66-7 (1931).—The fact that the colon bacilli do not give the indole reaction after some time is not a disadvantage as it shows that the pollution occurred long ago and that there is no danger of typhoid contagion.

The action of alkylresorcinols on phytopathogenic bacteria. C. STAPP *Angew. Bot.* 12, 275-289 (1930), *Rev. Applied Mycol.* 10, 46.—The effect of varying concns. of phenylethylresorcinol and of hexylresorcinol on phytopathogenic bacteria depends so greatly on the media in which the bacteria are grown that no conclusion as to the toxic action of these substances on phytopathogenic bacteria within the host can be drawn by their effect on the same organisms on artificial media. Where potato tubers were used as the media for *B. phytophthorus* concns. of 0.2% of hexylresorcinol and 0.3% of phenylethylresorcinol were necessary to inhibit the growth of the organism, whereas lab. tests showed that concns. of 0.03% of either killed the bacteria when free of media. Seven media and 7 bacteria were studied.

Gas-metal electrode potentials in sterile culture media for bacteria. ELTON M. BOYD AND GILFORD B. REED *Can. J. Research* 4, 64-68 (1931), *cf. C. A.* 23, 3411.—The work of French and Kahlenberg on gas-metal electrode potentials suggested that the potentials obtained at inert metal electrodes immersed in growing cultures of bacteria, and generally regarded as an oxidation-reduction phenomenon, might be of a similar type. This was tested by passing air, H_2 , N and CO_2 through sterile beef broth and observing the potentials developed at Pt, Au and Hg electrodes. It is shown that air causes a rise in potential at all 3 electrodes; H_2 causes a marked fall in potential at the Pt and a slight change at the Au and Hg electrodes, N and CO_2 cause little or no change in the potential at the 3 electrodes. These changes are in agreement with French and Kahlenberg's results except with the Hg electrode which is shown to react chemically with the broth.

The mechanism of bacterial action. J. H. QUASTEL *Trans. Faraday Soc.* 26, 853-84 (1930).—A summary of the work on the dehydrogenases of resting bacteria. The theory previously developed (*C. A.* 22, 249, 4137), of the localization of dehydrogenation reactions on sp. areas of cell surface, is outlined.

Increased bactericidal effect of inorganic compounds in the presence of x-rays (NORRIS) 7. Red colorations on salted bides and a comparison of the growth and survival of halophilic or salt-loving organisms and some ordinary organisms of dirt and putrefaction on media of varying salt concentrations (ROBERTSON) 29. The microbiological aspect of peat formation (THAYSEN) 8.

Handbuch der pathogenen Mikroorganismen. 3rd ed., edited by WILHELM KOLLE, RUDOLF KRAUS AND PAUL UILENHUTH. Band. VII, Tl. 2. Jena. G. Fischer. Berlin and Vienna. Urban & Schwarzenberg. Pp. 753-1752, M. 108. *cf. C. A.* 25, 983.

MACKIE, T. J., AND MCCARTNEY, J. E.: An Introduction to Practical Bacteriology: A Guide to Bacteriological Laboratory Work. 3rd ed., enlarged. New York: Wm. Wood & Co. 421 pp. \$3.50.

D—BOTANY

THOMAS G. PHILLIPS

Phototoxic substance from *Hypericum crispum*. Z. MÉLAS IOANNIDÉS. *Arch. Inst. Pasteur Hellén.* 2, 161-165 (1928).—In an attempt to isolate the phototoxic substance from *Hypericum crispum*, 0.16 g. of a red, amorphous substance, slightly sol. in water and giving a characteristic spectrum, was obtained from 100 g. of the dried plant.

The phototoxic pigment of *Hypericum crispum*. Zof. MÉLAS IOANNIDÉS. *Compt. rend. soc. biol.* 105, 349-51 (1930); *cf. preceding abstract*.—The isolated pigment of *Hypericum crispum* when inoculated into white rats photosensitized the animals in the same way that ingestion of the fresh plant does. The pigment was obtained in crystalline form by extg. the dried plants with 0.3% Na_2HPO_4 , and treating the filtrate with $AcOEt$. The 2 phases were sep'd. by satg. the aq. layer with Na_2CO_3 . The $AcOEt$ soln. of the pigment was purified by washing 3-4 times with sat'd. Na_2CO_3 soln., then with H_2O , the filtered soln. was conc'd. by distn. to 60 cc., and treated with 240 cc. of ether. The pigment was ext'd. by shaking with 0.3% Na_2HPO_4 . The pigment seps. in amorphous

form on acidification with KH_2PO_4 soln. Slow acidification, obtained by diffusion without agitation from KH_2PO_4 crystals, resulted in 0.05 g of brick red crystals. These are insol in CHCl_3 , CS_2 and C_6H_6 , very slightly sol in ether, fairly sol in MeOH , EtOH and AcOH and extremely sol in acetone and pyridine. All the solns are strongly fluorescent, and are bright red in color. The pigment is pptd in a brick red colloidal form by the addn of Na_2CO_3 or as a grayish green ppt on the addn of NaOH . It is also pptd by dil acids which seem easily to alter it. The following absorption spectra data were obtained on a soln of 0.1 g of pigment in 5000 cc of alc: 4 mm layer, band 1,599-588-586, band 2 553-515, 20 mm layer, band 1,605-587-575, band 2,558-547-545, band 3 513-549, 10-mm layer, band 1,612-573-570, band 2 562-530-528, band 3,515-503 (progressive absorption at the extreme violet of the spectrum, 485).

B. C. BAUNSTETTER

Wood-pentosan. I. K. NISHIDA AND T. TAKAGI. *Cellulose Int.* (Tokyo) 6, 170 3(1930). Abstracts, 32-3.—Woodmeals from a no. of different coniferous and deciduous trees were extd with 5% NaOH soln, the ext was treated with Fehling soln in order to sep pentosan as the Cu compd, the latter was decomposed with HCl , and the yield and compn of the pentosan were detd. The pentosan content (from the Cu compd) of the deciduous woods examd is found to be characteristic, while coniferous woods contain much smaller quantities of pentosan. Frequently the wood pentosan was converted into 2 different modifications, deciduous woods contain both the α and β varieties, while only the α type is present in coniferous wood. The compn of the β pentosan (% xylan etc.) from the deciduous woods examd is given; that of the α -variety will be given in part II.

B. C. A.

Micro-inoculation of diatoms without carapace. E. BACHRACH AND MME. PILLET. *Compt rend* 190, 1442-3(1930).—No SiO_2 was detected in the hyaline body of the diatom *Nitzschia* without carapace. The ash contained Ca.

B. C. A.

Differentiation between good and poor germinative capacity of seeds by chemical means. H. O. L'ACEN. *Dis. Agr.-Chem. Inst. Univ. Breslau* 1929.—The germinative activity of seeds is closely related to their dehydrogenase activity. Of numerous colorimetric methods for the detn of the latter, the most satisfactory is that using m-dinitrobenzene. Comparative tests should be made on the same day with freshly ground material.

B. C. A.

Action of a sudden withdrawal of nutrient salt solution from corn seedlings. G. BERTHOLD. *Z. Botan* 23, 94-131(1930).—Corn seedlings, grown for a time in nutrient salt solns, were transferred to distd water, and progressive anatomical and physiol changes were observed.

R. C. BUXELL

The nature of chemodimetrical stimulation and the differences in swelling produced by L-asparagine. HANS FITTING. *Z. Botan* 23, 328-60(1930).—Plasma streaming in leaf sections was studied and the stimulating effect of dil solns. of asparagine noted. These dil asparagine solns readily become inactivated as the result of bacterial action.

R. C. BUXELL

Chlorophyll changes in barley. HARALD NILSSON. *Arkiv. Kemi. Mineral. Geol* 10A, No 12, 34 pp (1930).—No essential differences in activity were shown between the normal green chlorophyll in homo- or heterozygotes or the chlorophyll-deficient white homozygotes in barley seeds. The O_2 utilization of the intact green embryo and of certain organs as the leaves, roots and endosperm is greater than that of the white parts. No difference between the homo- and heterozygotes could be detd in this respect. The earlier results which showed no relationship between O_2 utilization and catalase activity are confirmed. Higher O_2 consumption occurred in all cases with chlorophyll-contg leaves than with chlorophyll deficient ones. The presence of KNO_3 was without effect on the O_2 utilization. Only slightly lower ash content was found in the chlorophyll deficient leaves than in the green ones.

H. J. DEUEL, JR.

Plant proteases. XVI. The activation of animal and plant proteases by glutathione. W. GRASSMANN, O. V. SCHÖNBERG AND H. EISELER. *Z. physiol. Chem.* 194, 124-36(1931), cf. *C. A.* 24, 868.—Zöökinae, the natural activator of animal proteases, has recently been shown by Waldschmidt-Leitz (*C. A.* 24, 5316) to be identical with glutathione. A similar identification of the phyto kinase of yeast is now reported. The activator of yeast proteinase is rapidly liberated by autolysis in the presence of a cell poison, the greater part passing into soln during the first hr. The liberation of proteolytic enzyme, on the other hand, is hardly appreciable during the 1st hr and requires 24 hrs for completion. The sepn of enzyme and activator is therefore a comparatively simple matter. The 1st fraction, which has no enzymic activity, contains the glutathione of the yeast and has the property of activating the inactive enzyme to the full activity attainable by treatment with HCN . Boiled yeast ext shows the

same behavior. Fractionation of the crude activator soln., according to the procedure for purification of glutathione, yields fractions in which the activating capacity and the glutathione content are strictly parallel. The activity of the crude activator soln. is therefore due to its glutathione content. Since cysteine and glutathione must be in their reduced form to activate papain and cathepsin, it would appear that the natural SII system effects an intimate connection between respiration and proteolysis. Hydrolytic processes are commonly associated with anaerobic conditions, while synthetic processes are performed by the cell during active respiration. In the functioning cell, which is abundantly supplied with O, the glutathione would be present in its disulfide form, and hence the cathepsin for the most part would be in its inactive form. The cell would thus be protected from autolysis. However, there is reason to believe that the activation phenomena cannot be attributed exclusively to glutathione, since yeast proteinase undergoes self activation under conditions which preclude the intervention of glutathione.

Plant microchemistry. L. ROSENTHALER. *Pharm. Zentralhalle* 72, 81-5(1931).—A. W. DEX
An address. W. O. B.

Growth and absorption by oats in relation to varying concentrations of calcium and sodium in the nutrient solution. C. PHILLIPSON. *Svensk Bot. Tid.* 24, 185-202 (1930), *Physiol. Abstracts* 15, 263.—The relation of dry wt. of culture to composition of the nutrient soln. is shown graphically for 2 periods of growth: (a) 1-3 weeks, involving 6 harvestings, (b) 2 months and for mature plants and seeds. The dry-wt.-time curves for the first 3 weeks follow V. H. Blackman's formula. N content and ash content follow the same course as the dry wt., Ca content increases more rapidly. Changes in Ca utilization are involved in the antagonism between Ca and Na. It is probable that these changes are due partly to alteration in permeability of cell membranes. G. G.

The proteins and vitamins in the embryo of the soy bean. SAJIRO ONTOMO. *Abstracts from Rept. Central Lab. S. Manchuria Railway Co.* 1929, 15-7.—The embryo constitutes 2.09% of the soy bean. It contains moisture 4.47, crude protein 42.21, protein 35.76, crude fiber 2.90, ash 3.45, crude fat 9.40, non-nitrogenous ext. 37.42 and pentose 5.7%. Of the total N, 78.8% is sol. in H₂O, 8.61% in NaCl soln. and 5.22% in NaOH soln.; 78.68% of the N is present as globulin, 6.01% as albumin, 0.12% as proteose and 5.56% as nonprotein. On the basis of dry matter N is distributed as follows: total N 7.09, sol. in 20% HCl 0.60, amide 0.60, humin 0.21, arginine 1.32, histidine 0.23, lysine 0.55, cystine 0.09, amino 4.80%. The embryo is richer in vitamins A and B than is the endosperm. V. F. HARRINGTON

Absorption and utilization of potassium by plants. R. P. BARTHOLOMEW AND GEORGE JANSSEN. *Proc. Assoc. Southern Agr. Workers, 31st Ann. Convention*, 212-5 (1930).—The min. amt. of K which it is necessary to maintain in soln. to produce optimum growth varies with different plants. Alfalfa and Hubbard clover required a min. K concn. approximating 0.5 p. p. m., cow peas, oats, soy beans and cotton 2.0 p. p. m., and Sudan grass 3.0 p. p. m. In all expts. good growth was made from a concn. of only 0.5 p. p. m. Plants have the ability of absorbing relatively large amts. of K very rapidly, and the amt. absorbed depends upon the amt. in soln. and not upon the needs of the plant, large amts. being taken up when no visible growth is being made. The plant quickly adjusts itself to changes in the concn. of K. Although plants absorb K very rapidly symptoms of K starvation in the later stages of growth are infrequent because the plant is able to transfer the K from the older to the embryonic regions and re-utilize it for growth processes. K. D. JACOB

Isotopes of potassium: their association with plant life. J. G. F. DRYCE. *Chem. News* 142, 33-5(1931); cf. *C. A.* 24, 1885.—Addnl. exptl. work confirms D's earlier statement that KCl from the ashes of potato stalks contains more K⁴¹ than is present in ordinary KCl. In non-living plant tissue there is no accumulation of the higher isotope. LOUISE KELLRY

Spiko disease of *Dodonaea viscosa*. B. N. SASTRI AND N. NARAYANA. *J. Indian Inst. Sci.* 13A, 147-52(1930); cf. *C. A.* 25, 315.—Diseased tissues of *Dodonaea viscosa*, as compared with healthy tissues, showed a higher nitrate, nitrite, NII₂ and amide content, a tendency toward accumulation of sugar and starch, a higher diastatic activity and a deficiency in Ca. *Zizyphus oenoplia*, which suffers from a similar disease, also shows Ca deficiency. LOUISE KELLRY

Proteins of Indian foodstuffs. III. Globulins of Bengal gram (*Cicer arietinum*, Linn.) and horso gram (*Dolichos biflorus*). NUGHALLI NARAYANA. *J. Indian Inst. Sci.* 13A, 153-5(1930); cf. *C. A.* 25, 211.—A knowledge of the compn. of Bengal gram and horso gram is important because of the relation of these pulses to human and animal nutrition, resp. The globulins in these grams were isolated and were found to

contain sufficient quantities of arginine, tyrosine and lysine, but to be deficient in cystine and tryptophan. The arginine content of the Bengal gram was much higher than that usually found in globulins of other Indian pulses. LOUISE KELLAY

Physiological studies on *Drosera*. II. The effect of quinine and atoxyl on pepsin. KUNIO OKAHARA *Science Repts Tôhoku Imp. Univ* [4], 5, 739-55(1930), cf C A 25, 983 —In the course of studies made by O to det whether *Drosera* enzyme and pepsin are identical, he detd the effect of quinine HCl and atoxyl on pepsin. Digestion expts were carried out at 39°, and the extent of digestion was detd by the nephelometric method, a 40% soln of sulfosalicylic acid was used to cause turbidity. Variations in the method of mixing the solns, and in the concn of the solns of poison used caused the digestion to be accelerated, repressed or unchanged. LOUISE KELLEY

Composition and properties of certain red and yellow plant pigments. LEONORA N. BILGER *Bull Basic Sci Research* 3, 37-46(1931) —The pigment (I) extd from Japan red peppers is shown to differ in chem compn from *lycopson* and the pigment from paprika. Ultimate analyses for I gave C 74.69 and 74.69%, H 9.95 and 10.26% O (by difference) 15.36 and 15.05%. The crystals were deep red, m p 167-9° (when taken immediately after purification), 120-1° (after slight fading) and 50-60° (after complete fading), they were moderately sol in EtOH and MeOH, sol in C_2H_5 and ether, easily sol in $CHCl_3$ and CS_2 , nearly insol in petroleum ether and insol in water. They were purified by pptn from ether by adding petroleum ether and chilling and by recrystn from CS_2 . I shows no rotation. The dry crystals of *lycopson*, *carotene* (II), *capsanthin* and I are readily bleached when exposed to air and light. Alc. solns of I and II may be kept in stoppered glass test tubes (air present) for weeks in daylight at room temp without noticeable fading and for months in a refrigerator. Alc. and $CHCl_3$ solns of I and II were exposed in water-cooled quartz tubes to ultra-violet light and in cellophane cover cells to soft x rays with and without air present. In general, alc. and $CHCl_3$ solns faded much more rapidly than CS_2 solns. In the presence of air, alc. and $CHCl_3$ solns completely lost their colors after 3 hrs' exposures to ultra-violet light. CS_2 solns required 9 hrs. for complete fading or 26 hrs. when air was excluded. In all cases of exposure to ultra violet light, a very small amt. of a brown resinous material formed on the walls of the tubes. When the filtered solns were evapd to small vols and cooled sparkling colorless crystals sep'd. Fading of the pigment was also brought about with soft x rays but in a much shorter time than with ultra-violet light. After irradiation with ultra-violet light neither I nor II showed bands in the visible region; I but not II showed bands in the ultra violet. Twenty five photographs of absorption spectra of these pigments are appended. J A KENNEDY

Preliminary study upon the nucleo-cytoplasmic ratio in plant tissues. F. E. HUELVIN *Australian J Exptl Biol Med Sci* 6, 69-64(1929) —The ratio nucleic acid N/(coagulable protein N nucleic acid N) found for wheat plants 29, 46 and 57 days old was 0.0341, 0.0124 and 0.0040, resp. C G KING

Variability of rice varieties in resistance to toxic action of potassium chlorate and its practical significance. MORIMASA YAMASAKI *J. Imp. Agr. Expt Sta (Japan)* 1, 1-22(1929) —Rice seedlings grown in tubes of $KClO_3$ soln (0.005-2.0%) were found to vary in susceptibility to toxic action. Upland rice was more resistant than lowland varieties, and among strains of upland rice, a direct correlation between drought resistance and resistance to toxicity was shown. In all varieties there was a similar relation between length of growing season and resistance. Injury from $KClO_3$ was shown by a dark brown discoloration of the leaf veins. Conditions of growth previous to placing seedlings in $KClO_3$ soln also affected resistance, those grown under upland conditions and exposed to the sun were less resistant than those previously grown in a watered nursery and shaded. Resistance to toxicity decreased with the age of the seedlings. Differences in resistance to toxicity are attributed to varietal differences in permeability of root cells, not to differences in osmotic pressure of cell sap. K. KITSUTA

Variation and correlation among varieties of wheat and barley in regard to resistance to the toxic action of potassium chlorate. MORIMASA YAMASAKI *J. Imp. Agr. Expt Sta* 2, 139-62(1929) —See C. A. 24, 4576. K. KITSUTA

Variations in the unsaponifiable lipides and in the lipide phosphorus of *Stenogrammatocystis nigra* as a function of the mineral composition of the liquid culture medium. CHARLES PONTILLON *Compt rend* 191, 1367-9(1930) —The mineral compn of the nutritive medium has a considerable influence upon the lipide metabolism of *Stenogrammatocystis nigra*. B S LEVINE

Frangularoside, a new rhamnoside of the freshly dried bark of black alder. M. BRIDEL AND C. CHARAUX *Compt rend* 191, 1374-6(1930); cf C. A. 25, 934 —A

method for the prepn and purification of the rhamnoside is given and its properties are described. The use of freshly dried bark is essential. B. S. LEVINE

The distribution of nitrogen in plant extracts that contain a high proportion of nitrate nitrogen. ALBERT C CHINNALL AND B J. MILLER *J. Biol Chem* 90, 189-96 (1931)—Exts of plant products which contain a high percentage of nitrate N show certain abnormalities in the distribution of N, and it has been necessary to revise the methods employed for the partition of the N into the now commonly accepted groupings. Reduction of nitrate was found to occur during HCl hydrolysis (cf Vickery and Pueber, *C A* 25, 1552-3), this reduction was promoted by the presence in the ext. of an easily oxidizable substance which apparently was nitrogenous in character. A particular sample of rye grass (*Lolium perenne*) was found to contain 12.7% of its total N as nitrate N, an amt equal to 4.4% of the dry wt. of the leaf. Only the tobacco leaf has previously been shown to contain such a relatively enormous amt of nitrate. A. P. LOTHROP

The fermentable principle of the tubers of the asphodel. II. C NEYRON. *Bull sci. pharmacol* 38, 38-51(1931); cf *C A* 25, 1274.—The rhizomes and tubers contain the glucoside *asphodeloholoside* which was obtained as a Ba(OH)₂ compd sol in water and pptd. by dil EtOH. It is partly decomposed by CO₂. The glucoside itself cannot be crystd, it is extremely sol in water and EtOH. The mol wt is about 700, the rotatory value is between -18° and -19°. After hydrolysis, the rotation is -67°; the sugars formed are levulose and an aldose, probably glucose. It is easily fermented, especially by *Aspergillus niger*, and only partly by sucrose. A. E MEYER

Presence of methyl mercaptan in the leaves of *Lasianthus laevigatus* Bl., *L. lucidus* Bl., *L. purpureus* Bl., *L. stercorarius* Bl., and *L. bracteolatus* Miq. D R. KOOLHAAS *Biochem. Z.* 230, 446-50(1931)—A method is discussed for the detn. of MeSH in the steam distillate from leaves. S. MORGULIS

Chemistry of citric acid formation by molds. IV. Transformation of saccharic acid. K. BERNHAUER, H. SIEDENAUER AND H. TSCHINKEL. *Biochem. Z.* 230, 460-74(1931); cf. *C. A.* 23, 1151.—Expts with 26 strains of mold fail to give any evidence for the formation of saccharic acid from glucose or gluconic acid, or for the formation of citric acid from saccharic acid or its salts. There is thus no evidence that saccharic acid is an intermediate product in citric acid formation. A no. of the molds transformed saccharic acid to oxalic acid. S. MORGULIS

Characterization of *Aspergillus niger* strains. III. Further propagation experiments and comparison of different mold strains. K. BERNHAUER, F. DUDA AND H. SIEDENAUER. *Biochem. Z.* 230, 475-83(1931); cf. *C. A.* 23, 2741.—The different strains are differentiated by their ability to produce citric acid and to form gluconic acid. S. MORGULIS

Isotopes and living organisms. V. I. VERNADSKII. *Compt. rend.* 192, 131-3 (1931).—The work of Loring and Druce (*C. A.* 24, 1885) indicates that K extd from the potato has an at. wt. of 40.4-40.59 and is more radioactive than the lower isotope of K. These results are contested by Lowry (*C. A.* 25, 314) who extd the K of cotton and of wheat and obtained an at. wt. of 39.09-39.31. V. prepd compds of K, Fe, Mg, Zn, Ca, Si and S from living organisms with a view to det. the at. wt. of each element. These data would throw some light on the hypothesis that living organisms sep. the isotopes of elements. This hypothesis is supported by the fact that N, C, H and other elements occur in living organisms in isotope mixts., and that certain elements (such as Ca, K, Fe and Mg) present in the vital cycle tend to remain there. N. M. NAYLOR

The liberation of iodine from the gland cells of *Bonnemaisonia asparagoides* by the action of ultra-violet light. ROBERT LAMM *Compt. rend.* 191, 863-5(1930).—Irradiation of *Bonnemaisonia asparagoides* with ultra-violet light causes the liberation of I₂. It is possible that the algae are killed, and that the liberation of acids from the diffuse vacuoles causes this effect, but it is probable that it is a direct effect of the radiation which ionizes the free I₂ present in the gland cells. T. H. RIGER

Studies on the nature of rust resistance in wheat. V. Physiology of the host. W. F. HANNA. *Can. J. Research* 4, 131-47(1931).—Detns were made of catalase, diastase and oxidase activity; rate of respiration; and content of chlorophyll, xanthophyll and carotene in the leaves of 8 wheat varieties. With respect to their reactions to stem rust, these wheats vary from almost complete susceptibility in certain varieties to a high degree of resistance in others. Catalase activity increased as the plants approached maturity, whereas diastase activity decreased with increasing age. No significant differences were found in the oxidase activity or respiratory rate of the varieties. Little Club and the varieties of the vulgare group proved to be relatively rich in chlorophyll and the carotenoids. It is suggested that photosynthetic processes may

take place more rapidly in the cells of the varieties having a high content of these pigments, and thus furnish conditions suitable for the growth of the rust mycelium

J W SIMPLEY

The mechanism of growth in the cotton hair. F. T. PEIRCE *Trans Faraday Soc* 26, 809-13(1930); cf. *C A.* 23, 5381.—The deposition of sol carbohydrate as cellulose involves linkages in 3 directions, each one corresponding to a stage in the physiology of the plant. The theory of orientation of glucose residues in the cell wall is developed

K V THIMANN

Useful device for evaporating alcohol from plant extracts. F. E. GARDNER. *Plant Physiology* 5, 617-8(1930)—The device is attached to a Soxhlet extn battery. It consists of an air pump driven by a $\frac{1}{4}$ h-p. motor and connected to a piece of 0.75-in iron water pipe 8 ft long into which are fitted, at the proper intervals, 12 curved pieces of $\frac{1}{16}$ in Cu tubing about 7 in long which lead into the necks of the flasks. The time necessary for the evapn of 100 cc. 95% alc is 15-20 min and for 80% alc considerably less than 1 hr. A diagram of the app is given

WALTER THOMAS

Some organic acids of wheat plants. L. K. NELSON AND HEINRICH HASSELBERG. *J Am Chem Soc* 53, 1040-3(1931)—In an investigation of the org acids in growing wheat plants the aq ext was found to contain (% calcd on the basis of green plants) malonic, by fractionation of esters, 0.005, acetic, by Li_2O extn of acids, 0.040, malic, by fractionation of esters, 0.056, by polarization, 0.064; citric, by fractionation of esters, 0.016, by pentabromoacetone, 0.019, oxalic, trace. The residue after water extn contained 0.02% oxalic acid (0.11% on dry basis)

C J WEST

Influence of neutral red on yeast respiration. MAX GEIGER HUBER. *Proc Acad Sci Amsterdam* 33, 1059-68(1930)—Manometric measurements of the O_2 consumption at 20° showed that the respiration of bakers' yeast is independent of the pH between 5 and 8.1. When neutral red is added 1/10,000, respiration in acid medium quickly rises 20-30% and after an hour falls to 80% of the original. In alk media it falls immediately. The older the cells, the more pronounced are these changes

J J WILLAMAN

Changes occurring in stored alcohol plant extracts. JAMES E. WEBSTER. *Science* 73, 77-8(1931)—Plant samples preserved in 80% EtOH + excess CaCO_3 were found to have pH values of 4.6-6.0 and titers of 3-8. Celery exts in 80% EtOH for 9 months doubled their NH_3 content at pH 4.2 and quadrupled it at pH 8+. There was always increase in NH_3 in exts of grapes, celery, lettuce and spinach. W. urges that this method of preservation be studied by other investigators

J J WILLAMAN

Changes in the composition of oranges during ripening. P. R. V. D. R. COPEMAN. *Trans Roy Soc S Africa* 19, 107-67(1931)—The growth and compn of oranges were followed for 104 days, fruits from both unsprayed trees and from those sprayed with Pb arsenate were used. Most of the factors studied fitted rather well this modified form of equation for autocatalytic reactions $\log(x - c)/(a - x) = K(t - t_1)$, where c is the value of the factor at the beginning, a is the final value, $K = k(a - c)$ (k is the velocity constant of the change) and t_1 is the time at which x has a value of $a + c/2$. The factors studied were wt. of fruit, skin and pulp, alc insol residue, sol solids, acidity, sucrose, glucose, fructose, ash, N, P, K, Ca, Mg, Fe. The proportion of P, both in the juice and in the ash, decreased during the period of ripening, and storage of P must occur in the early stages of formation of the fruit. The K content of the juice and of the ash increased during ripening and the relatively large proportion of K in the mineral constituents of the juice indicates clearly the importance of this factor in the development of the fruit. The results indicate that the growth of the fruit is dependent upon a continual supply of K. It would appear that K is closely connected with the formation of carbohydrates in the plant, and it is believed that this factor is also assoc. with the transference of the carbohydrates. The K content of the juice of the sprayed oranges is greater than that of the unsprayed oranges, and the proportion of sugar in the total solids was also higher in this case. The Ca and Fe contents of the juice decreased during ripening, while there was no significant rate of change in the content of Mg and no significant increase in N or in ash contents in the juice. Ca appears to resemble P in that it must be stored in the early stages of development of the fruit. P-N ratios in the case of both sprayed and unsprayed oranges are similar and show no significant difference, the mean values being 0.352 and 0.351, resp. The value of the ratio decreased in both series during ripening. On the other hand, the K-N ratio was slightly higher in the case of the sprayed oranges, the mean values for the 2 series being 1.83 and 1.67, resp. The value of this ratio remained more or less const. during ripening. During this period the wt. of the orange increased, and this was accompanied by an increase in the wt. of pulp, in the amt. of sol. solids and sugars in the juice and in the amts. of these constituents per fruit

At the same time the proportion of cell-wall material in the pulp and of the acid in the juice decreased. During the final stages of growth transpiration became dominant, with the result that the fruit lost wt. because of loss of water. At the same time there was a corresponding increase in the concn of sol solids in the juice. The acidity of the sprayed oranges was less than in the unsprayed oranges, while there was a greater content of cell wall material in the sprayed oranges. At the same time there was a slightly higher content of sugars in the juice of the sprayed oranges, and this appears to be due to the fact that the amt. of sucrose was increased. The rates of change of the various factors were not affected by the spraying, and it was concluded that the arsenate spray exerts an internal physiol. action on the trees, with the result that there is a selective oxidation of the respiratory materials in the fruit. As a consequence of this action the acid in the sprayed oranges forms a more important source of respiratory material than the cell wall material. The action must be caused by the absorption of arsenical compounds through the leaves of the trees. The solution of the problem depends upon the use of sprays other than arsenical. A bibliography is included. J. J. WILLAMAN

E—NUTRITION

PHILIP B. HAWK

Changes in the composition of blood in rabbits fed on raw and cooked soy beans. ERNEST TSO AND S. M. LING. *Proc. Soc. Exptl. Biol. Med.* 28, 219-20 (1930). —The feeding of raw or cooked soy beans or a diet of millet and cabbage had no effect on the concns. of uric acid, urea N and inorg. P in the blood of the rabbit. The blood cholesterol value was slightly higher on the soy bean diets. Uric acid was detd. by the Folin improved method (cf. *C. A.* 24, 2771). Twice the vol. of blood filtrate was used and the standard was reduced to 2 cc.

C. V. BAILEY

Effects of a hyperprotein and hyperpurine diet (pancreas) on the rat. P. RONDOVI. *Arch. ist. biochim. ital.* 2, 243-68 (1930). —On a diet of pancreas, either alone or mixed with flour, rats lose wt., especially in the former case, while the kidney wt. increases both absolutely and relatively. Other effects of such diet include slight tendency to leucocytosis; diffuse phenomena of swelling in the parenchymal organs; more marked vital colorability with trypan blue; an acidotic condition, increased thermal reactivity compared with that caused by subcutaneous injection of peptone or suspension of uric acid, increased disposition to inflammation.

B. C. A.

Ergosterol and antirachitic vitamin. M. COMEL. *Arch. ist. biochim. ital.* 2, 281-306 (1930). —When administered in excessive doses, irradiated ergosterol maintains its efficacy as a Ca fixing and antirachitic remedy but gives rise to toxic phenomena. The toxicity is displayed mainly by pptn. of lime in the viscera and by a characteristic atrophy of their cellular elements, and is probably due to the ergosterol itself rather than to admixed substances. This toxicity throws doubt on the identity of irradiated ergosterol with the lime fixing vitamin, but the irradiated ergosterol may represent the active constituent of the vitamin, the latter being formed by union of the ergosterol with a protective lipid.

B. C. A.

Effect of x-rays on ergosterol. M. SUMI. *Bull. Inst. Phys. Chem. Research* (Tokyo) 8, 640-4 (1929). —Vitamin D, formed by irradiation of ergosterol with ultra-violet light, is destroyed by 4 hrs' exposure to x rays. The spectrum of ergosterol and activated ergosterol extends further into the ultra-violet region after irradiation with x rays.

B. C. A.

Biological investigations with milk changed by the removal of fats and the addition of carbohydrates. L. RANDOM AND R. LECOCQ. *Compt. rend. soc. biol.* 102, 523-31 (1929); cf. *C. A.* 25, 729. —Pigeons were fed daily with a mixt. prep'd from 8 g. of agar, 2 g. of filter paper and 90 g. of dried milk or dried milk from which fats had been removed. The animals fed on the fat free diet died in 30-50 days. In another series, the diet consisted of 50 g. of dried milk (fat free or normal), 40 g. of carbohydrate and 10 g. of roughage. With the fat free diet, the animals fed with lactose, sucrose, maltose, dextrose and dextrin died in 15-25, 50-70, 70-100 days, 3-4 months and not less than 5 months, resp. With the normal milk diet, only the animals fed with lactose died (50-80 days).

B. C. A.

Dependence of the sugar curves of the blood on acid and alkaline diets with subcutaneous injection of pilocarpine. E. T. MINNIK-BOGDANOVA. *Russian J. Physiol.* 13, 151-6 (1930). —An acid or an alk. diet influences the sugar of the blood but little, the sugar content and its max. and min. values being approx. the same in the 2 cases. With either diet, subcutaneous injection of pilocarpine produces neither a definite increase nor

a definite lowering of the sugar content of the blood, but the divergences from the mean value are more marked than those of most of the control expts. B. C. A.

Variations of the catalase curve of blood after introduction of pilocarpine on an acid or alkaline diet. S. A. POVOZINSKAYA. *Russ. J. Physiol.* 13, 183-92 (1930).—The curves representing the catalase content of the blood at various times after subcutaneous injection of pilocarpine are of a uniform type for children receiving an alk. but not an acidic diet; where a mixed diet is taken, the curves are of a mixed type. The catalase-time curves for children on an alk. diet not receiving pilocarpine are not of a uniform type. The catalase curves vary irregularly for the same individual according to the day and hour at which the blood is taken. No parallelism is found between blood catalase and hemoglobin contents. B. C. A.

Variations in the chloride and water contents of striated muscle, liver and kidney during acute experimental scurvy. LUCIE RANDOIN and ANDRÉ MICHAUX. *Compt. rend.* 192, 108-10 (1931), cf. *C. A.* 24, 4535, 25, 1559.—In guinea pigs deprived of the antiscorbutic vitamin the Cl and water contents of the liver and kidney increase only slightly above normal, but the Cl content of the muscle increases considerably. This is not dependent upon the Cl content of the diet, as the same effect was noted in guinea pigs receiving diets contg. no NaCl, 0.5% NaCl and 1.5% NaCl. RACHEL BROWN.

Observations and measurements on vitamin A. P. KARRER, B. EULER, H. von EULER, H. HELLSTÄD and M. RYDBOM. *Arkiv Kemi, Mineral. Geol.* 10B, No. 12, 0 pp.—A study is made of the metabolic variations of carotene and its relationship to vitamin A. Carotene was found to be concentrated in the blood plasma or serum. Almost none was present in the red blood cells even in those bloods which contained large amounts of this pigment. In 50 samples of normal human blood sera the av. was 0.8 Lovibond units per 10 cc., although one value was 4.5 units. With herbivorous animals, much higher values were obtained. With ox serum the content was 17 units and with cow sera taken in Sept. the av. was 14 Lovibond units (11-16). A seasonal variation occurred in these animals, the highest concns being obtained in the fall after a summer of fresh pasturage while much lower amounts were obtained after the usual winter feed. While it is admitted that carotene is not the growth-promoting factor identical with vitamin A, it is suggested that this latter factor may be a hydrogenated carotene which has been formed in the liver through the action of carotenase. A spectrometric study of these pigments in hysterearin and in the distillates of cod liver oil and shark liver oil definitely proves that a growth promoting factor independent from carotene exists. A very rapid deterioration of these substances occurs, as indicated by the rapid disappearance of the absorption bands. Other sol. substances having a different spectrometric picture were obtained from germinating barley and *Fucus vesiculosus*. H. J. D. JA.

The ether-soluble substances of polished rice. YOSHIMIZU SAHARU. *Bull. Inst. Phys. Chem. Research (Tokyo)* 9, 882-90 (Abstracts 84-5 (in English) published with *Sci. Papers Inst. Phys. Chem. Research (Tokyo)* 14, No. 271-3 (1930)).—A further substantiation of earlier expts. (*C. A.* 23, 1937; 24, 3264) in which it was demonstrated that polyneuritis in birds is not to be traced to toxic compds in the alc-sol. fraction of polished rice as claimed by Teru Uchi (*C. A.* 24, 1146). The ether sol. portion consists of palmitic, oleic and linoleic acids, phytosterol and other oily nonsaponifiable matter. The injection of this ext. was not followed by any toxic symptoms in pigeons or in mice, nor was any difference noted in the effect of feeding alc-extd. polished rice or that washed only with water. In either case the diet must be supplemented with vitamin B to prevent the onset of polyneuritis. H. J. DEUEL, JR.

The vitamin content of marine plankton. J. C. DAUMOND and E. R. GUNTHER. *Nature* 126, 394 (1930).—Phytoplankton gave a pos. test for vitamin A in 20 mg. doses and the SbCl₅ and absorption spectrum tests were likewise pos. These were neg. for zooplankton. The line test for vitamin D was only doubtfully pos. with phytoplankton but unmistakably pos. with zooplankton, while the x ray gave neg. results with both species. H. J. DEUEL, JR.

The relation of vitamins to carcinogenesis. WARO NAKAHARA. *Sci. Papers Inst. Phys. Chem. Research (Tokyo)* 15, 32-6 (1930); *Chem. News* 142, 86-7 (1931).—In opposition to Fujimaki, Kimura, Wada and Shimada (cf. *C. A.* 21, 1836), N. finds no direct relationship between vitamin-A deficiency and cancer in rats. Overdosage with vitamin D was likewise ineffective in producing this condition. H. J. DEUEL, JR.

Problems ofomenclature and dosage in experiments with Vigantol. M. SCHMIDT-MANN. *Klin. Wochschr.* 9, 1629-30 (1930). H. EAGLE.

The use of an acid diet in the treatment of bronchial asthma. KURT TEIFENSEN. *Klin. Wochschr.* 9, 1675-8 (1930); cf. *C. A.* 24, 2798.—Untreated cases of bronchial asthma show a decreased excretion of acid and NII, during an attack. An acid diet,

with addn of NH_4Cl , $\text{NH}_4\text{H}_2\text{PO}_4$ and "Silicalcium," leads to a significant improvement in the asthmatic condition. Asthmatics kept in a secluded room show an increased excretion of acid and a decreased no. of attacks. Definite improvement, however, was obtained only when an acid diet was given. Alkalosis is therefore an important symptom in bronchial asthma, the correction of which affects the whole clinical picture.

H. EAGLE

A new principle in the biological demonstration of irradiated ergosterol. JOSEF WARKANY *Klin. Wochschr.* 9, 2152-3(1930)—Rabbits which usually excrete no phosphate, and are given 30-100 mg of irradiated ergosterol in their food, show large quantities of phosphate in the urine in 48 hrs. A normal phosphate excretion can be prevented by daily CaCO_3 in the food.

H. EAGLE

The fate of oxidizing enzymes in foods in the human digestive canal. I. BOAS *Klin. Wochschr.* 9, 2295-6(1930)—Using bananas (which are particularly rich in catalase, oxidase and peroxidase) as a test meal, B. finds that all 3 are inactivated if the stomach is acid (to Congo red), but that this inactivation is reversible, the enzymes being reactivated by the neutralization of the mixt. Oxidase is the most sensitive to acid. As shown by the passage of a duodenal sound, as well as by expts. on a dog with fistula of the ileum, the enzymes are reactivated when the food reaches the intestines. Normal feces contain no oxidase. If pus is present, oxidase may appear, if blood is present, peroxidase may appear. Catalase is normally present. Normal urine contains no peroxidase or catalase, but they may appear in pathol. conditions. Oxidase may be present, but it is thermostable and is probably a metallic catalyst.

H. EAGLE

The active principle of liver in pernicious anemia. KURT G. STERN *Klin. Wochschr.* 10, 172-5(1931)—A summary of the recent work.

H. EAGLE

The chemical nature of vitamin D. ETSUO TAKAMIYA. *J. Dept. Agr. Kyushu Imp. Univ.* 3, 1-27(1930)—A repetition and extension of studies previously reported (cf. C. A. 24, 3523, 25, 123, 536). New work includes a comparison between the action of O_2 and of ultra-violet light upon cholesterol and olive oil with respect to photoactivity, color reactions, and m. p. of the cholesterol. The behavior of the ozonized materials was similar to that of those exposed to ultra-violet light. Ozonized ergosterol was found to be antirachitic when fed to rats made rachitic on McCollum's diet 3f43, the potency rising to a max. and then falling off with increasing time of exposure to O_3 . The relationship of these exptl. results to the chem. nature of vitamin D is discussed, and it is tentatively concluded that the antirachitic portion of the ozonized ergosterol is ergosterol monoozonide Δ_1 , which is possibly identical with vitamin D.

S. B. FOSTER

Influence of radium emanation on the transformation of ergosterol to vitamin D. J. MAISON, W. MUND, Y. POURBAIX AND A. CASTILLE. *Compt. rend. soc. biol.* 103, 534-6(1930), *Physiol. Abstracts* 15, 258—Ergosterol in soln. in paraffin irradiated by Ra rays in a closed receptacle undergoes a definite chem. transformation. The product obtained is endowed with antirachitic properties.

G. G.

Reducing power of tissues in animals with low cystine diet. N. C. LACLAU AND A. D. MARENZI. *Rev. soc. argentina biol.* 5, 768-73(1929); *Physiol. Abstracts* 15, 233.—The reducing power is less than in the controls. Differences are greater in liver than in muscle. There is a certain relation between reducing power and glutathione content of tissues.

G. G.

Experimental production of bladder stones. S. PERLMANN AND W. WEBER. *Munch. med. Wochschr.* 77, 680-1(1930); *Bull. Hyg.* 5, 699—Twenty-three % of the rats fed on a vitamin A-deficient diet had phosphate stones in the bladder. Of 100 rats fed on an unsuitable but not deficient diet, half died in 3 months; of the remaining 50, 7 showed gravel at the end of 80 days. The small stones consisted of urates and resembled those of human beings. All normally fed rats were free from stones. Vitamin-A deficiency is not the only factor, but it is an important one. Europeans have a predominance of urate stones, while oxalate stones predominate among the Japanese. In the case of children, stones are probably due to improper feeding, but in adults anatomical factors, infection and disordered metabolism are the probable causes.

GEORGE R. GREENBANK

The nutritive value of soy-bean cake and purified soy-bean cake. SAJIRO OHTOMO. *Abstracts from Rept. Central Lab. S. Manchuria Railway Co.* 1929, 10-2—Feeding and artificial digestion expts. show that the alc-extd. cake is less nutritive.

V. F. H.

The occurrence of vitamins A and B in soy-bean oil and cake. SAJIRO OHTOMO. *Abstracts from Rept. Central Lab. S. Manchuria Railway Co.* 1929, 13-5—Rat and pigeon feeding expts. indicate that soy-bean cake and oil contain some vitamin A. This is contrary to McCollum, *et al.* (C. A. 10, 2902).

V. F. HARRINGTON

Studies on experimental rickets. III. The isolation of ergosterol from brewer's

yeast and the activation of ergosterol by ultra-violet light. SENCIN IZUME, YOSHINOBU YOSHIMARU AND ISAO KONATSUNARA. *Abstract from Rept. Central Lab. S. Manchuria Railway Co.* 1929, 18-21.—The ergosterol, m. 160-1° and having $[\alpha]_D^{25} = -122^\circ$ in CHCl_3 , after irradiation cured rickets in doses of 0.001 mg. per day. Rays of 311-290 μ are most effective in imparting antirachitic power. Too long exposure destroys this power. V. F. HARRINGTON

Recent developments in vitamin research. W. D. SALMON. *Proc. Assoc. Southern Agr. Workers, 31st Ann. Convention*, 293-02(1930).—In general, diets in the South are rather low in vitamins A and C and dangerously low in what for the present may be called the vitamin G fraction, including the pellagra preventive factor. They are also deficient in Ca and in vitamin D. There is little doubt that an abundance of sunlight is the only thing that prevents rickets from being more prevalent in this region. K. D. JACOB

Dietary standards for pregnancy. CALLIE COONS. *Proc. Assoc. Southern Agr. Workers, 31st Ann. Convention*, 293-5(1930).—A study was made of the retention of N, Ca and P by 9 pregnant women at different stages of fetal development. There was a distinct tendency for N retention to increase with an increase in intake and vice versa, and N was stored much in excess of the probable fetal demand. Storage of Ca was highest in late pregnancy, as was also the curve of probable fetal demand for Ca. Only 3 of the 9 women were storing as much Ca as the probable fetal demand, and many were storing much less. Assocd. with low Ca retention were instances of maternal tooth destruction and of poor bone formation as shown by x ray examn. of the 8-day old baby. Retention of P was even more const. than intakes, the same woman storing as much in the middle months as in late pregnancy. Also the rate of storage was well above the fetal demand in the middle months. K. D. JACOB

Antirachitic activation of ergosterol by soft x-rays. ELIZABETH SNELOW AND JOHN R. LOONBOUROW. *Bull. Basic Sci. Research* 3, 47-64(1931).—Theoretical considerations show that any radiation of shorter wave length than about 3000 or 3100 A. U. should be capable of activating ergosterol. This applies to radiations throughout the x-ray and γ ray regions. The extent of activation should not be appreciable, however, except in portions of the spectrum in which the absorption is of consequence. Calens show that the greatest x ray absorption of ergosterol lies in the extreme soft x ray region and that appreciable activation should be expected in or near this region. Irradiation expts. with soft x rays are described, and absorption spectra and animal expts. indicative of the activation of ergosterol by soft x rays are recorded. The effect of x rays on the solvent might be more pronounced than on ergosterol or cholesterol with solvents containing elements of higher at. wt. than H, C or O, and consideration must be given to the possibility of glass containers giving fluorescent ultra-violet radiations in the presence of x rays. The effect of soft x-rays in activating ergosterol may be a primary or secondary one. The literature is reviewed and photographs and x ray pictures are given. J. A. KENNEDY

Modern problems of nutrition. AUGUSTIN D. MARENZI. *Ann. farm. biogum. Supplemento* 2, 1-35(1931).—A comprehensive review of the subject along the following lines: I. Complete diets. Factors of importance: equilibria (a) caloric, (b) N, (c) mineral, (d) vitamin. II. The importance of the nitrogenous substances: (a) protein minimum, (b) practical protein minimum, (c) high- vs. low-protein diets. III. The nutritive value of proteins. Supplemental value. Its relation to the amino acid content. (a) Methods for the detn. of nutritive values of a substance, (b) the nutritive value of proteins from the viewpoint of maintenance and (c) of growth, (d) supplemental value of proteins, (e) the relation between biol. value of proteins and their chem. constitution. IV. The relation between amino acids and maintenance, and the growth of animals. Methods of study. General results. Conclusions. There is no bibliography. B. S. LEVINE

Antioxidants and the autoxidation of fats. H. A. MATILL. *J. Biol. Chem.* 90, 141-51(1931).—To secure information on the chem. nature of the antioxygenic substances that are found in natural oils and that prevent autoxidative destruction of fat-sol vitamins, a series of hydroxy aromatic compds. was tested for their capacity, when used in a quantity of 0.02%, to prolong the thermal oxidation induction period of a standard mixt. of lard and cod liver oil. The antioxygenic capacity of phenols resides in 2 OH groups in the *o*- or *p*-configuration; when these are in the *m*-position the compd. is inactive. The OH groups are ineffective unless attached directly to the ring; fully hydroxylated inositol is ineffective. In the naphthols 1 OH group is sufficient, and, in keeping with its accepted behavior, α -naphthol has the character of an *o*-compd. and is

much more effective, as an antioxidant, than β -naphthol; quinone is effective and β -naphthoquinone is even more so, but the α -form is entirely inactive. A no. of sterols of animal origin and sitosterol from wheat, corn and lettuce were all inactive. The existence of pro- and anti-oxygenic substances among the unsaponifiable constituents of natural fats and oils suggests that some of these may be concerned with the physiol. action of the fat-sol vitamins and methods for their segregation from the sterols are under investigation. It is suggested that in the prepn and manipulation of easily autoxidizable substances, the presence of traces of antioxidant will prevent undesirable oxidative changes. A bibliography of 42 references is given. A. P. LOTHROP

The effect of nitrous acid upon the components of the vitamin-B complex. H. C. SHERMAN AND M. L. WHITSITT. *J. Biol. Chem.* 90, 153-60 (1931).—The "protein-free" milk prepn of Osborne and Mendel was used as a suitable material contg. both vitamins B (B_1) and G (B_2). The solns. were treated with HNO_2 in 2 ways: in one the gas was drawn by aspiration into the vitamin soln. and in the other the HNO_2 was generated directly in the vitamin soln. by adding $NaNO_2$ to it and dropping in an equiv. amt. of 1.1 HCl. After 14 hrs. in each case the dissolved gases were removed under reduced pressure and the liquid was then neutralized with 0.2 M NaOH. Treatment by the aspiration method had little, if any, influence upon vitamin B but a large diminution of potency occurred when the HNO_2 was generated in the soln. These results suggest that vitamin B may be a nitrogenous base but of such a structure as to be more resistant to HNO_2 than are typical primary amines. No destruction of vitamin G was brought about by aspiration of HNO_2 through the soln., but partial destruction by the 2nd method was indicated by feeding tests when the data were taken for an 8-week but not for a 4-week exptl. period. It seems probable from these expts and other unpublished observations that vitamin G_1 (B_2) is stable toward HNO_2 but that there is more or less destruction of one of the newer and not yet clearly defined factors making up the vitamin-G complex and that the reaction involved is likely one of oxidation rather than deamination. A. P. LOTHROP

Studies on alimentation and nutrition. IV. Means and measures of food economy. C. FOA. *Arch. sci. biol.* (Italy) 15, 445-64 (1930).—This is a review and is part of a series of investigations promoted by the Italian National Research Council. The subjects discussed are bread, distribution of food resources between man and domestic animals, utilization of leaves, utilization of waste and residues from slaughter houses, refrigeration and food preservation, utilization of refuse as fertilizer, savings in kitchen and table practice and methods for evaluating the food requirements of man.

PETER MASUCCI

The chemistry of nutrition. FILIPPO BOTTAZZI. *Atti III congresso naz. chim. pura applicata* 1930, 51-64.—The latest developments in the biochemistry of carbohydrates, proteins, enzymes and vitamins are reviewed and discussed. The study of metabolism by thermodynamic methods alone is considered insufficient. A detailed knowledge concerning the chemistry of nutritive substances is essential, in addition, to knowing their energy content, before the biol. value of the substances can be ascertained.

PETER MASUCCI

The phosphate esterification and the phosphatase action in B-avitaminosis. J. BODNÁK AND ANTON KARELL. *Biochem. Z.* 230, 233-7 (1931).—The muscles and livers of pigeons suffering from vitamin B deficiency esterify H_2PO_4 77 and 146%, resp., more than the organs from healthy birds. No difference was found between the normal and vitamin B deficient pigeons so far as the phosphatase activity of the muscles and livers is concerned.

S. MORGULIS

Intermediate carbohydrate metabolism. VI. Sugar assimilation, phosphorus and water metabolism during prolonged intravenous injections of glucose, fructose and galactose. M. WIERNICHOVSKI, W. PIESKOW AND E. OWSIANY. *Biochem. Z.* 230, 146-72 (1931); cf. C. A. 22, 3232.—The total reduction of deproteinized blood from 3 well nourished dogs, in fasting condition, ranged between 91 and 104 mg. % of which 54-67 mg. % was due to glucose and 34-40 mg. % was due to residual reduction. In the continuous intravenous injection of glucose for 3 hrs. at the rate of 2 g. per hr. the blood sugar during the first hr. rose on the av. of 300 mg. % but fell during the next 2 hrs. This biphasic reaction was not observed in the other hexoses. Two hrs. after the injection the blood sugar was lowered below the initial level. During the expt. 10.5% of the injected glucose was eliminated in the urine, the glucosuria curve like the glucemia curve being also biphasic. The rest of the injected sugar was utilized. Two hrs. after the injection there was no more glucosuria. When fructose was injected, the fructose concn. of the blood rose gradually but was on the av. lower than the glucose concn. in parallel expts. However, of the total amt. of injected fructose 10.5% was lost in the

urine, the rest having been assimilated. Two hrs after the injection the fructose disappeared from the blood, and only traces were found in the urine. With galactose injections the blood galactose concn. increased very rapidly, becoming 800 mg % in the middle of the third hr, and not until 9 hrs after the injection did the galactose disappear from the blood and urine. Only about 30% of the galactose was utilized, the rest being excreted in the urine. Under the influence of insulin the diaphasic nature of the reaction after the injection of glucose disappeared, since the glucemia in the first hr was much diminished and the assimilation definitely increased. The insulin had practically no effect on the fructosemia or on the utilization of fructose, but it did greatly stimulate the rate of removal of galactose from the blood and improved its assimilation to such an extent that only 42% was excreted in the urine. The injection of fructose or galactose of itself does not affect the glucose concn. of the blood, but under the influence of insulin the glucose concn. diminishes while the concn. of the fructose or galactose actually rises. The "residual reduction" remains unchanged when either glucose or fructose was injected. In these expts no relation was found between the blood and urine P values. With a const. hexose injection the blood became most acid. with galactose and least acid with fructose. The quantity of urine corresponded to the amt. of excreted hexose.

VII. Lactic acid production under the influence of prolonged intravenous injection of glucose, fructose and galactose. M. WIRZUCHOWSKI AND M. LANIEWSKI. *Ibid* 173-85. —The lactic acid concn. of the blood of fasting dogs varied from 7 to 25 mg %, the hourly excretion in the urine ranging from 1.4 to 5.1 mg. In expts with a continuous intravenous injection for 3 hrs at the rate of 2 g. per hr. the blood lactic acid curve for each hexose presented rather definite characteristics. The lactic acid excretion ran parallel to the changes in the blood concn. Most lactic acid was formed from fructose and occasionally from galactose, and least from glucose. After glucose injections about 0.1 g. lactic acid was excreted, after galactose injections 0.1-0.3 g. and after fructose injections 0.3-0.6 g. more than before the injections. With insulin the lactic acid excretion was increased with all 3 sugars. The lactic acid formation is responsible for the acidosis during the sugar injection, the rise in lactic acid concn. of the blood being more or less correlated with the increase in concn. of the hexose.

VIII. Respiratory metabolism of intravenously injected glucose, fructose and galactose. M. WIRZUCHOWSKI. *Ibid* 187-224. —During the expts with continuous intravenous hexose injections the sv increase in the nonprotein R. Q. was for glucose 0.23 without insulin and 0.26 with insulin, for fructose 0.24 and 0.30, and for galactose 0.13 and 0.18, resp. The sp dynamic action of the glucose was 9.7%, of fructose 13.8% and of galactose 17.3% of the caloric content of the assimilated part of the hexose. The greatest rise in heat occurred between the second and third hrs and was 51% for glucose, 63% for fructose and 32% for galactose. In the expts with insulin 2 periods were distinguished: hyperglucemic, while sugar and insulin were being injected, and hypoglucemic when nothing was being injected. In the first period the increase in heat production was for glucose 21%, fructose 33% and galactose 23% greater than in corresponding expts without insulin. The greatest increase in heat was 57% for glucose, 85% for fructose and 32% for galactose. During the hypoglucemic period, 4-5 hrs after the injection, the heat production was increased 36%. The largest amt. of sugar which the dogs were able to oxidize was 0.5 g. per kg., and of the total assimilated sugar 34% of the glucose, 41% of fructose and 56% of the galactose underwent oxidation. During the hyperglucemic period under the influence of insulin the oxidation was increased 15% for glucose, 24% for fructose and 87% for galactose. On the av 26.7% of the cal of the different hexoses was utilized in the sp dynamic increase of heat production, and when 1 mol. sugar was oxidized 235 cal were liberated. Control expts. in which physiol. NaCl soln was injected showed only an insignificant rise in heat production. During the 2nd and 3rd hrs of the injection when the transformation of the hexose was at a max. the heat production and the increase in R. Q. were inversely proportional to the concn. of the injected hexose in the blood. Although a parallelism is not always observed between the blood lactic acid curve and the heat production curve, the extent of lactic acid formation from 1 g. of assimilated hexose seems to be correlated with the sp dynamic action and oxidation in that these are greater for the sugar from which the organism produces more lactic acid. S. M.

The iodine content of the bile and thyroid gland in cattle under the influence of the seasonal changes in feeding. G. PREIFER. *Biochem. Z.* 230, 290-8 (1931). —An increase in the av. wt. of the thyroid glands of cattle of various ages was observed from the winter months to the summer months. It has not yet been detd. whether there is also a decline in wt. from the summer to the winter months. The I content rises both with age and with the gland wt. and reaches a max. value in cattle 5-6 years old. The change from the I-poor winter fodder to the I-rich summer fodder has no

recognizable influence upon the I of the thyroid glands. The bile secretion was found to increase with age. The I excretion in the bile during the summer is 120% higher than in winter, which shows that the I of vegetable matter is well absorbed by the animals. The I content of the feces cannot serve as an accurate guide to a study of I absorption.

S. MOROGLIS

Determination of value of vitamin D preparations. II. MARTIN SCHIEBLICH. *Biochem. Z.* 230, 312-9 (1931), cf. C. A. 23, 4971.—In testing the potency of vitamin D preps, the healing dose does not correspond generally to the preventive dose, the former always being higher and the difference between the 2 being very variable. Detn. of the healing dose gave practically the same results either by the line test or by the x-ray exam. In such tests there is no evidence that there is any danger of error due to eating of feces, but there is apparently danger from animals licking each other. S. M.

The excretion of creatinine bodies a function of the amount of the endogenous nitrogen discharged. EMILE P. TRASSOIN, R. DONNET, P. DANMANVILLE and G. MOUROT. *Compt. rend.* 191, 1473-5 (1930).—This series of expts. was carried out to det. whether the behavior of creatinine bodies in metabolism can be attributed to the N of exogenous or of endogenous origin. The diet contained glucides for the energy requirement and was adequate in mineral salts and vitamins but contained no protein. Inanition expts. were carried out with pigs of 15-20 kg. and upon rats. Intoxication expts. after complete inanition were carried out with pigs for the addn. of BrO₃H and of phlorhizin to the diet, and with rabbits for the addn. of H₂PO₄. The results of the expts. indicate that (1) the excretion of creatinine remains const. though that of endogenous N varies greatly; (2) the excretion of creatine follows that of total N; it rises with a rise of excretion of endogenous N as influenced by the diet. The creatinine bodies behave in endogenous N metabolism exactly as in exogenous metabolism. N. M. NAYLOR

Irradiated helisterol and its antirachitic power. G. MOURIQUAND and A. LEULIER. *Presse méd.* 37, 1261-2 (1929); *Bull. soc. hyg. aliment.* 18, 200 (1930).—Irradiated purified cholesterol extd. from snails was found to possess practically no antirachitic power. On the other hand, cholesterol contg. the cholesteric lipoids of snails (helisterol) was found to be exceptionally active, but only after irradiation. One to two drops of a 0.1% helisterol soln. was found sufficient to prevent rickets in rats and to facilitate fixation of Ca in the case of an animal fed exclusively on Pappenheimer's diet No. 85.

A. PAPINEAU-COUTURE

Clinical investigations on the antirachitic action of irradiated helisterol. G. MOURIQUAND and A. LEULIER. *Paris méd.* 19, 409 (1930); *Bull. soc. hyg. aliment.* 18, 204; cf. preceding abstract.—Clinical observations showed that irradiated helisterol has an antirachitic power comparable to that of ergosterol. This power reveals itself in the various forms of human rickets, particularly in cases of so-called "florid" rickets and in those in which the decalcification lesions do not extend to the whole diaphysis.

A. PAPINEAU-COUTURE

Vitamin A in serum and liver. BERTH V. EULER. *Swensk Kem. Tids.* 42, 302-4 (1930) (in German).—There is no vitamin A in the subcutaneous fat of rabbits. There is probably no vitamin A in the serum of rabbits and sheep. The SbCl₅ color reaction is indefinite and is probably due to something other than caratene. A. R. ROSE

The digestion of keratin, especially the horny material of birds' feathers, by fowls and mammals. ERNST MANGOLD and JOSEF DUBISKI. *Wiss. Arch. Landw., Abt. B., Tierernähr., Tierzucht* 4, 200-21 (1930).—The secondary quills of white goose feathers were fed to owls, cats, dogs and rats. Examin. of the feces showed some mech. disintegration in the case of the cat and owl, but balance expts. failed to show digestion in any of the expts. Rats on a diet deficient in cystine showed no growth that could be attributed to cystine furnished by the feathers ingested. W. GORDON ROSE

An investigation of the vitamin A content of liver meal. ANTON SOHNER. *Wiss. Arch. Landw., Abt. B., Tierernähr., Tierzucht* 4, 231-90 (1930).—The vitamin A potency of liver meal was investigated to det. whether it is a suitable source of vitamin A for milk cows and young animals during the winter months. Rats were fed a diet prep'd. as directed by v. Euler and v. Euler, *Zur Methodik der Vitaminprüfung* vol. 5, which was deficient only in vitamin A. Liver meal was added to the diet when noticeable loss in wt. occurred. Xerophthalmia was cured and normal growth effected by 0.1 g. per day. This was equiv. to 5-10 mg. cod-liver oil. Since several of the rats lost wt. without showing any indications of xerophthalmia, other criteria, such as loss in wt., are necessary in vitamin A studies. W. GORDON ROSE

The biological values of wheat, barley and rye. ISTVÁN WEISER and ARTHUR ZATTSCHIRK. *Wiss. Arch. Landw., Abt. B., Tierernähr., Tierzucht* 4, 401-6 (1930).—Goats were fed diets of hay plus wheat, hay plus barley and hay plus rye, and weighed at

Intervals The expt. showed no essential difference in the biol values of the 3 grains.
W. GORDON ROSE

Antirachitic vitamin D. Announcement of a standard by the medical research council. ANON *Lancet* 1930, II, 503-4—A quantity of ergosterol from yeast, purified and dried by methods carefully recorded, was subjected, in an accurately made alc. soln., to irradiation with the rays from a Hg arc, the phys. details of the procedure being measured and recorded with the greatest practicable accuracy. The irradiation products were carefully freed from alc., dissolved in pure, unadulterated olive oil and the vol was adjusted to a known concn., in terms of ergosterol used. Such solns. retain their activity unchanged for 2 years if maintained at or below 0°.

Effect of added iron on the hematopoietic properties of dried milk. A. L. BACHA RACH *Lancet* 1930, II, 957-8—The recognized low hematopoietic properties of milk solids can be materially enhanced by drying appropriate quantities of an Fe salt in the milk.
F. B. SEIBERT

The toxicity of vitamin D. J. B. DEGUIN *Lancet* 1930, II, 983-5, cf. C. A. 24, 5058—The pathol. effects in animals of overdosage of vitamin D are: a tendency to increased serum Ca and blood loorg. P; a very greatly increased excretion of Ca and P, especially by way of the kidneys, involving a reduction of Ca and P retention in the body, an accumulation of Ca salts in the tissues and loss in body wt. To estg. the dosage and toxicity a large no. of factors must be taken into account. Standardization is urged.
F. B. SEIBERT

The constitution of irradiated ergosterol. A. CASTILLE *Bull. acad. roy. méd. Belg.* 10, 319-25 (1930)—A study of the absorption spectra of aliphatic geometrical isomers shows a remarkable similarity between the *cis*- and *trans*-compds. Since the same similarity is found between the spectra of ergosterol B and Isoergosterol B as well as between those of dihydroergosterol A and dihydroisoergosterol A, a simple *cis*-*trans* isomerism due to the mobility of the secondary alc. groups is assumed. Comparison of the absorption spectra of ergosterol A (I) and of its irradiation product ("vitamin D") points to the fact that the biol. activation is due to the gradual formation of a geometrical isomer isoergosterol A (II) that is identical with the active principle of vitamin D. Prolonged irradiation leads to further rearrangement into another biologically inactive modification, this transformation begins before all of I has been converted into II.
G. TOENIES

The quantitative estimation of the glutathione content in normal and pathological tissues. III. Relationship between the development and the glutathione content in chick embryos. TOKUO KAWAYA *Nogoya J. Med. Sci.* 5, 1-5 (1930); cf. C. A. 24, 4057—Expts. made on chick embryos show that the glutathione content increases gradually during the first half of development, reaches a max. in 12-14 days, and then gradually decreases. The liver, brain, heart, muscle and eye contain the largest percentage of the glutathione of the body and are richest in the order named. IV. The glutathione content in various organs of pigeons fed on polished [rice] and subjected to starvation, and also of starved rabbits. *Ibid.* 6-17.—Pigeons fed on polished rice showed a decreased glutathione content in liver, heart and muscle proportional to the no. of days of the expt. The adrenal glands, spleen and blood showed an increase. Starvation produced the same result, but to a greater degree. Symptoms of polyneuritis appeared after 15-18 days in pigeons fed on polished rice and later in those fed on half polished rice. V. Relationship between glutathione content and neurotomy in muscles of pigeons fed on polished and unpolished rice. *Ibid.* 18-23.—Pigeons were subjected to neurotomy of the sciatic nerve of the right leg, the left serving as control. The wt. of the right leg did not increase as rapidly as that of the left and the glutathione content of the muscle subjected to neurotomy showed an increase proportional to the no. of days elapsed since the operation. This was true for pigeons fed on polished rice as well as unpolished rice, but in the control muscles, the glutathione content decreased more rapidly in the former case.
A. WIKEN

Ergosterol and some of its derivatives (DE BOE) 10.

DRUMMOND, J. C., AND HILDITCH, T. P. The Relative Values of Cod Liver Oils from Various Sources. London: H. M. Stationery Office. 135 pp. 1s. Reviewed in *Chemistry & Industry* 50, 160 (1931)

F—PHYSIOLOGY

F. K. MARSHALL, JR.

Excitant substance in the central nervous system. L. HABERLANDT. *Arch. ges. Physiol. (Pflügers)* 223, 171-9 (1929)—The excitant action of emulsions of the cerebrum

or midbrain when injected into frogs is due to a substance which is sol. in water, but not in alc or Et_2O , it is dialyzable, and is decomposed when heated. B. C. A.

Existence, origin and significance of the so-called lutein colloid. A. ALBANES. *Arch 1st biochim ital* 2, 269-80(1930) —Discussion of previous work, together with new exptl results, indicates that the ruby-red test is not sp. for the lutein colloidal substances, that the presence of this colloid is not sufficient for the diagnosis of a gravid lutein body, and that the so-called colloid encountered in either the menstrual or gravid lutein body is merely a metamorphic blood deriv., which is accompanied by involutive, atrophic phenomena of the lutein body itself. B. C. A.

Amylase content of the milk of various animals. R. SCHENK. *Arch wiss prakt. Tierheilk* 58, 375-84(1928) —Human milk contains a considerable quantity of amylase, colostrum free milk of the cow, horse, dog, cat and guinea pig contains only traces, while neither milk nor colostrum of the goat contains amylase. The markedly higher amylase content of the colostrum of the other animals, falling with the progress of lactation, indicates that the enzyme originates in the colostrum. B. C. A.

Glutathione content of human placenta and denervated gastrocnemius of the toad. J. DI LEO LIRA. *Compt rend soc biol* 102, 469-70(1929) —The former contains 0.029-0.038%, the latter 0.014-0.018% (normal), 0.012-0.017% (denervated) of glutathione. B. C. A.

Lipogenic and antilipogenic action of hormones as cause of obesity and leanness. A. LUBLIN. *Klin Wochschr* 8, 2276-8(1929) —Hyperfunction of lipogenic hormones (e.g., insulin) causes obesity and of antilipogenic hormones (e.g., adrenaline, thyroxine, pituitary hormone) causes leanness. Intravenous injection of thyroxine (1 mg.) causes a transient rise in blood sugar (e.g., 0.110-0.130 g. per 100 cc.). B. C. A.

Hydremia in alimentary hyperglucemia. M. N. KALLINIKOVA. *Russ J Physiol* 13, 193-7(1930) —Detn. of hemoglobin and water in the blood of children, taken at various times after ingestion of sugar, show that hydremia is not assocd with hyperglucemia. B. C. A.

Cholesterolemia in alimentary hyperglucemia. M. N. KALLINIKOVA AND G. D. ONRAZTZOY. *Russ J Physiol* 13, 198-203(1930) —In children alimentary hyperglucemia is assocd with cholesterolemia, which attains its max. value 15-30 min. after ingestion of sugar. No general parallelism is found between the blood-cholesterol time curves and the dextrose-time curves. B. C. A.

Blood lipase in alimentary hyperglucemia. E. T. MINKER-BOGDANOVA AND G. D. ONRAZTZOY. *Russ J. Physiol* 13, 204-9(1930) —Max. variations of the lipase-time curves of blood following the ingestion by children of large quantities of sugar coincide with the maxima of the dextrose and fat content-time curves. The types of lipase-time curves found for different individuals correspond either with the dextrose-time curves or with their mirror images. B. C. A.

Absorption of water, 0.2% hydrochloric acid, gastric juice, ethyl alcohol and various chlorides by the dog's isolated stomach. R. O. FAIRBLEY. *Russ J. Physiol* 13, 224-36(1930) —From 22 to 31% of water introduced into the isolated small stomach of the dog is absorbed within 60 min. Addn. of HCl to 0.2% or of NaCl to 0.9% reduces the percentage absorption to about 17%. Natural gastric juice in some cases undergoes resorption, assocd with diminution in acidity. The abs. but not the relative amt. of chlorides resorbed from hypertonic solns. increases with concn. Absorption of alc. from 10 and 20% solns. is complete within 60 min., and amounts to 90% in the case of 40% alc. solns. The velocity of absorption of chlorides and of alc. is greatest during the 30 min. following introduction. B. C. A.

The variations of the specific gravity of the plasma and the means available for altering it. NORMAN S. MOORE AND HAROLD J. STEWART. *J. Clin. Investigation* 9, 423-39(1930), cf. C. A. 24, 3828 —For comparable results the same anticoagulant must be used, heparin being employed in these expts. Normal specimens of blood show a remarkable constancy of the sp. gr. of plasma and of plasma proteins, provided the samples are taken at the same time each day. In patients suffering from heart disease who were in water balance the max. daily variation in sp. gr. was 0.0004. In 2 normal individuals the variations over a period of 14 months were 0.0015 and 0.0011. Arterial blood showed lower values than venous. Intravenous injection of normal saline soln. into dogs lowered the plasma sp. gr. and protein. Administration of 1 l. of water by mouth to human beings had slight effect. After hemorrhage a decline in these values was followed by a return to normal after 4 days. The injection of gum acacia in glucose soln. slightly increased the sp. gr., but lowered the proteins by diln. J. B. BROWN.

Carbon dioxide and oxygen tensions of the mixed venous blood of man at rest. DICKINSON W. RICHARDS, JR., AND MARJORIE L. STRAUSS. *J. Clin. Investigation* 9,

475-532(1930)—Certain aspects of the process of equilibration of lung gases with incoming venous blood are discussed. Equil can be regularly established for CO_2 with certain high O mixts and for both CO_2 and O with certain low O mixts. J B B

Studies on the Arneth count. XVII. The effect of alterations of the serum-calcium level on the count. DAVID R. CLIMENKO. *Quart J. Exptl Physiol* 20, 369-77(1930), cf C A. 24, 5351.—Irradiated milk taken by humans, colloidal Ca orally administered to rabbits, and physiol saline soln intravenously injected into rabbits produced no change in the Arneth count or the serum Ca level. Colloidal Ca injected intramuscularly into rabbits and CaCl_2 intravenously raised the serum Ca level and produced a regenerative deflection of the count. Intravenous injections of hypotonic solns produced an effect on the serum Ca level but caused a regenerative deflection of the count. RACHEL BROWN

The reformation of muscle glycogen destroyed by work. G. DABOIS. *Compt rend soc biol* 103, 546-8(1930)—Contraction of the paw muscles of cats (anesthetized or decerebrated) was effected by elec stimulation of the sciatic nerves. The integrity of the vagi nerves is essential for the reformation of muscle glycogen after work; the left vagus alone is sufficient. The glycogen lost after section of the cervical medulla, vagotomy or pancreatectomy is fully restored by the injection of 30 units of insulin. B. C. BAUNSTETTER

The effect of the injection of synthetic thyroxine on the carbon and nitrogen in urine. GEORGES FONTÈS AND LUCIEN THIVOLLER. *Compt rend. soc. biol* 105, 554-9(1930)—Two adult male dogs of the same wt and age were kept on a const diet throughout the expt: after a month, the vol of the urine, as well as its C and N contents, was recorded daily for 3 weeks (I), the same detns were made over a period of 2 weeks in which 1 mg of thyroxine was injected daily (II), over a period of 2 weeks in which 2 mg of thyroxine was injected daily (III) and lastly, over a period of 2 weeks in which no injections were made (IV). Variations of wt. over each period were recorded. The av amts of C and N per day in g were I, 3.60, 5.04, II, 4.12, 5.66, III, 4.15, 5.72, and IV, 3.62, 4.90. The C/N ratios for the 4 periods were 0.71, 0.72, 0.71 and 0.73. The vol of urine was not altered by thyroxine treatment. The av wt changes were I, +150; II, -450, III, -400, and IV, +250. B. C. BAUNSTETTER

The effect of thyroidectomy with and without injections of thyroxine on the carbon and nitrogen in urine. GEORGES FONTÈS AND LUCIEN THIVOLLER. *Compt rend. soc. biol* 105, 559-61(1930), cf preceding abstr.—The same diet and technic were used in this expt. A 15-kg dog was used. The data obtained support the conclusion that thyroidectomy alters the C/N ratio of the urine, but that injections of thyroxine do not. The thyroid gland exerts functions other than the elaboration of thyroxine. B. C. BAUNSTETTER

The effect of delivery on the protein equilibrium of the serum. CH. ACHARD, M. BARIÉTY AND A. COUDONIS. *Compt rend. soc biol* 103, 1064-6(1930)—The total protein, the albumin and the globulin of the sera of 9 women at the time of delivery and 10 days after delivery were detd. The total protein and albumin are slightly less than normal and can be accounted for by the hydremia of pregnancy. In 6 of the cases, the total protein was decreased 10 days after delivery, in 5 cases the albumin was decreased, while in 5 cases the globulin was increased. B. C. BAUNSTETTER

The organic bases and amino acids in normal human urine. SAKISABURO WADA. *Acta. Schol. Med. Univ. Imp. Kioto* 13, 187-99(1930)—From 300 l of normal human urine, the following compds were isolated: leucine, alanine, L-proline, D-proline, isoleucine, valine, phenylalanine, aspartic acid (?), glutamic acid, hippuric acid, creatine, creatinine, histidine, methylguanidine, betaine (?), choline, methyl pyridylammonium hydroxide, lysine, xanthine, L-methylxanthine (?), adenine, hypoxanthine, heteroxanthine and p-xanthine. H. J. DREUL, JR.

Cortin, the vital hormone of the adrenal cortex. FRANK A. HASTMAN. *Endocrinology* 14, 229-32(1930)—The injection of cortin (an ext. prepd from the adrenal cortex) not only prevented the development of the usual symptoms following double adrenalectomy but allowed the animals (cats) to survive indefinitely (170 days in the longest expt) as well as to increase in wt. Its potency was also demonstrated in bringing about the recovery of 2 animals in acute adrenal insufficiency. The blood urea is lowered from the high level occurring in adrenal deficiency. A relationship was shown to exist between the rate of growth of adrenalectomized rats and cortin injection. H. J. DREUL, JR.

The effects of growth-promoting and gonad-stimulating principles of the anterior lobe of the pituitary on basal gaseous metabolism in the rat. MILTON O. LEE AND JULES GAGNON. *Endocrinology* 14, 233-42(1930)—The subcutaneous administration

of the gonad-stimulating principle of the anterior lobe of the pituitary daily for periods of 8 to 24 days caused no significant change in the basal metabolism in rats. Identical results were obtained on exts. of the fresh gland and on that isolated from the urine of pregnant women. On the other hand, the intraperitoneal injection of these preps was effective in reducing the metabolic rate of 10 of the 18 animals used after several daily injections. The lowered metabolism continued 6 to 14 days after the treatment had stopped.

H J DEUEL, JR.

The metabolism of the isolated frog heart. A. J. CLARK, C. P. STEWART AND R. GADDIE. *Proc Roy Soc Edinburgh* 50, 297-303(1930).—Expts. were made on isolated frog heart perfused with Ringer's soln. either with or without glucose. Even in the presence of serum, glucose and insulin the carbohydrate metabolism represents only a fraction of the total metabolism. The protein catabolism can account for somewhat over 50% of the total O consumption. There is no evidence that this organ uses fat as detd. by the amt. present. The heart does not synthesize glycogen from protein. It is possible that an unknown factor essential for carbohydrate metabolism is absent in these artificial expts. although the ease with which the heart utilizes non carbohydrates sources of energy suggests that the metabolism of this foodstuff is not a primary process which produces muscle contraction. These results are at variance with those obtained earlier on dog heart.

H J DEUEL, JR.

Physiological chemistry of the blood in the tropics. W. BORCHARDT. *Arch Schiffs-Tropen-Hyg.* 34, 608-12(1930).

H EAGLE

The iodine content of the normal thyroid of albino rats. G. SANKARAN. *Indian J. Med. Research* 18, 563-70(1930).—Thyroids average an I content of 0.018%, independent of sex and age. Thyroid weight is proportional to body weight.

H. E.

The significance of the parathyroid hormone for the regulation of the calcium economy. CHRISTIAN BOMSKOV. *Klin. Wochschr.* 9, 2065-6(1930).—A preliminary report.

H EAGLE

The water-binding relationships of child blood plasma. J. JOCHIMS. *Klin Wochschr.* 9, 2115(1930).—Viscosity is a reliable index of the water-binding power of the plasma colloids. In child plasma it varies between 0.095 and 0.108. During hydremia in the child there is a parallel fall in plasma viscosity and protein. The sp. viscosity remains const., indicating that the excess water in the plasma is free and not bound by plasma colloids.

H EAGLE

Choline and choline esters in the blood. H. BORN. *Klin Wochschr.* 9, 2147-8(1930).—A protein-free ext. of human blood contg. choline, when injected into rabbits or cats, causes a very pronounced fall in blood pressure. Choline esters are known to be many thousand times as active in this respect as choline, and B. suggests that the choline in the blood (1.5-2 mg. %) is, in part at least, in the form of an ester with, e. g., AcOH. In essential hypertension there is a low blood choline.

H. EAGLE

Glucolysis and increase in acidity in tissue cultures, and analogous phenomena in the organism. A. A. KRONTOVSKII AND M. A. MAGATH. *Klin. Wochschr.* 9, 2165-8(1930).—The oxidation of glucose to lactic acid is the most important source of energy in tissue cultures. This type of metabolism may be characteristic of active regenerative processes even in the organism. The CO₂ formed results both from respiration and from the action of the lactic acid upon the NaHCO₃ of the medium. Changes in tissue cultures are analogous to changes observed in ischemia, inflammation, wound healing, regeneration, etc.

H. EAGLE

Milk secretion and menformone. S. E. DE JONGH AND ERNST LAQUEUR. *Klin Wochschr.* 9, 2344-6(1930).—The administration of menformone to male guinea pigs causes lactation. It is interesting to note that large quantities, 50 units twice daily for 4 weeks, cause a preliminary mammary hypertrophy; if this dose is then suddenly decreased, lactation begins within 2-3 days and continues for as long as 40 days if constantly decreasing quantities of menformone are injected. The authors point out the complete analogy to conditions in the pregnant woman.

H. EAGLE

The protein content of the urine of patients with normal kidney function. BRUNO MINZ. *Klin Wochschr.* 9, 2352-3(1930).—Using sulfosalicylic acid as a pptg. agent and solns. of casein as a nephelometric standard, M. finds that the urine of normal subjects invariably contains some protein, ranging from 0.14 to as high as 13 mg. %. The Esbach method can detect only 25 mg. %.

H. EAGLE

The theory of diuresis (urine excretion). HELLMUT MARX. *Klin Wochschr.* 9, 2384-8(1930).—The blood of a fasting dog contains a substance which inhibits the diuresis caused by water. The significance of this finding is discussed, particularly in the light of the fact that a pituitary ext. has a qualitatively similar effect.

H. EAGLE

Substances in human urine and diagnoses of pregnancy. M. ARON AND M. KLEIN

Compt rend soc biol 103, 702-4(1930); *Physiol. Abstracts* 15, 218-9 —The action of human urine on the thyroid glands of guinea pigs and rabbits demonstrates that the substance in the urine producing this effect is the hypophyseal hormone. This sp reaction is obtained with all urines, whether from men or women, but the urine of pregnancy only can affect the genital organs. This latter effect is due, not to the hypophyseal, but to a placental substance. G G

Pulmonary chemical sensibility studied by means of the isolated head. A. P. SÜSER AND J. PUCHER. *Compt rend soc biol* 103, 735-6(1930); *Physiol. Abstracts* 15, 220. G G

Secretion of cerebrospinal fluid. T. FALKIEWICZ. *Compt rend soc. biol* 103, 821-3(1930); *Physiol. Abstracts* 15, 253 —The secretion of the cerebrospinal fluid is stimulated by hypophyseal ext., pilocarpine and atropine; it is not increased by adrenaline, thyroxine, caffeine, AmH_2O_2 , theobromine, salicylate of Na, Br, I, luminal and hypertonic or hypotonic solns of NaCl. The hypophysis is the biol regulator of this secretion. The neg. results are opposed to the theory of transudation and confirm that of sp. secretion. G G

The storage of iodine in living tissues. W. WEBER. *Schweiz med Wochschr* 15, 237-9(1930), *Bull Hyg* 5, 912 —The thyroid is known to store appreciable amts of I, other glands such as the ovary and suprarenal store the element but in considerably smaller quantities. W fed rabbits KI and injected solns of the salt in the blood to det. whether the I content of the tissues was raised. Tissues such as the liver, heart, lungs and spleen were so low in I that the amt might be attributed to exptl error. The value for the kidneys was higher, but this might be attributed to the urine on the tubules. The urine contained considerable amts of I. After 24 hrs 50-60% of the I consumed had been excreted. GEORGE R. GREENBAUM

The content of glutathione of the blood in high altitudes and on exposure to natural Alpine sun. J. VON DESCHWANDER. *Strahlentherapie* 39, 278-82(1931) —D found that the glutathione content of the blood and that of the red blood cells are considerably increased at altitudes of 1300-5000 m above sea level. After one hour's irradiation with natural Alpine sun, the glutathione content of the blood and the quotient glutathione/erythrocytes decrease. This means that the glutathione content of the red blood cells decreases. Conclusion: The sp. respiration function of the blood and that of the red blood cells are increased at these altitudes, while by irradiation with natural Alpine sun this function is decreased. FREDERICK R. GREENBAUM

The minute volume and gaseous metabolism of the artificially perfused mammalian intestine. H. F. ROSS. *Arch ges Physiol (Pflügers)* 226, 181-9(1930) —The O_2 consumption of denervated mammalian intestine is 4.5 cc. per kg. per min., but varies considerably with time. The R. Q. is greater than unity. ARTHUR GROLLMAN

Ammonia formation and sugar utilization in artificially perfused mammalian intestine. H. F. ROSS. *Arch ges Physiol (Pflügers)* 226, 190-7(1930) —Intestines perfused by a Starling heart-lung prepn form appreciable quantities of NH_3 , presumably as a result of bacterial decompn. Glycocoli is resorbed by the intestine but is deaminized neither by the intestinal wall nor by the lung of the heart-lung prepn. Urea, although retained by the intestinal wall, is not converted to NH_3 . The sugar utilization of the intestine is 1.75-1.94 g. per kg. per hr., only 20% of which is oxidized. ARTHUR GROLLMAN

Effect of sleep, abstinence from sleep, and physical work on the blood chemistry. M. HOCHHAIM, J. MICHELSEN AND H. BECKER. *Arch ges Physiol (Pflügers)* 226, 244-54(1930) —Abstinence from sleep for a night resulted in an increase in the NaCl, total base, total CO_2 , and PO_2 of the blood. Physical exercise produced greater changes in the blood chemistry of subjects deprived of sleep, during the preceding night, than in normal rested subjects. ARTHUR GROLLMAN

The temperature coefficient of oxygen utilization by surviving organs. E. SCHMUTZLER. *Arch ges Physiol (Pflügers)* 226, 291-304(1930) —The O_2 consumption of isolated perfused limbs of the dog remained const. for 3-4 hrs. Cooling and subsequent rearming did not affect this O_2 consumption. The temp. coeff. for the O_2 consumption was const. and equal to 3% of the value at 35°, for each degree change in temp. A similar linear relationship also occurred in goose blood over the temp. range of 17-33°, the coeff. being 4% of the value at 35°. ARTHUR GROLLMAN

The effect of reduced air pressure on animals. III. The effect of varying degrees of reduced air pressure. W. LINZEL AND T. RAOFF. *Arch ges Physiol (Pflügers)* 226, 307-18(1930), cf C A 24, 4544 —White rats were kept for 3 weeks under reduced air pressures of varying degrees. The blood corpuscles increased at air tensions corresponding to 2000 and 4000 m. altitude, but little at higher altitudes. The total hemo-

globin and blood vols. increased in inverse proportion to the air pressure. The heart wts did not increase except at pressures corresponding to altitudes of 6000 meters or more. **ARTHUR GROLLMAN**

A chemical investigation in Victoria (Australia) of the blood of cattle and sheep. **JOAN H. NORRIS** *Australian J. Exptl. Biol. Med. Sci.* 7, 3-4 (1930), cf *C. A.* 24, 5809—The mean values found for inorg. P (mg./100 cc.), total P and Fe, resp., were 4.44, 10.13 and 60.54 for 26 steers, and 6.63, 12.59 and 57.92 for 22 cows. **C. G. KING**

A chemical study of the liver during inanition. **L. VAUOIN, M. JAVILLIER, H. ALLAIRE AND MAD. SCHIRMER** *Bull. soc. chim. biol.* 12, 894-902 (1930)—A dog receiving water only until its weight had dropped from 7 to 5.1 kg. showed the following % loss (dry wt.) in liver constituents compared to a control kept on a normal diet: glycogen and glucose 74.8, ash 33.7, insol. ash 48.1, ether ext. 31.8, cholesterol 24.4, unaccountable matter 10.5, glycerides 45.7, lecithin P 25.1, nucleoprotein P 16.5, N 18.3, total loss in dry wt. 27.2%. **C. G. KING**

The pituitary posterior hormone in fat metabolism. **W. RAAB** *Endocrinology* 14, 385-8 (1930)—Pituitrin promotes the absorption and destruction of circulating fat by the liver through a nervous pathway starting in the tuber cinereum and running through the cervical spinal cord and the abdominal splanchnic to the liver. Any disturbance of the pituitary mesencephalic system leads to obesity. Arterio-venous fat detns. suggest that pituitrin does not effect the fat deposit directly. There seems to be a relation between the pituitrin nervous mechanism and chem. heat regulation. **MARY E. LEAR**

Crystalline cholesterol in the post-retinal fluid. **ALBERTO J. SCHIJO AND LUIS N. PIZZORNO** *Ann. farm. bioquim.* 1, 160-6 (1930). **B. S. LEVINE**

Protein sugar in the blood plasma of the horse. **H. BIERRY** *Compt. rend.* 192, 240-1 (1931), cf *C. A.* 25, 1270—The protein sugar content of the blood plasma varies with the species and within the species with the individual. Furthermore, it is possible to provoke in the same individual important variations in the protein sugar content through the injection of adrenaline and insulin. In the blood plasma of the horse glucidic complexes combined with protein (protein sugar) constitute about 1/4 of the total sugar. The latter oscillates around 2 g. per l. of plasma. Marked individual variations occur. **B. S. LEVINE**

Combining the values of optic density and viscosity of a suspension for the determination of the number and volume of the dispersed elements. **G. ACHARD** *Compt. rend.* 192, 242-4 (1931)—A study of the relation between viscosity and no. of red blood cells on the one hand, and optic d. and no. of red blood cells on the other. An attempt was made by combining the viscosity and optic d. value to derive an empirical formula which would make possible the detn. of the vol. and the no. of the dispersed elements. **B. S. LEVINE**

A metabolic study of desiccated suprarenal medication in man. **R. G. HOSGINS AND F. H. SLEEPER** *Endocrinology* 14, 109-11 (1930)—Before the medication was begun records were made of a series of clinical lab. detns. The suprarenal substance was administered by mouth in dosages varying from 18 to 90 gr. daily for 64 to 108 days. The results are recorded in a table. From the viewpoint of lab. findings they are neg. Likewise no significant mental changes are observable. The general deviations observed were such as are usual in any medically selected group. **B. S. LEVINE**

Purification of the male hormone by high-vacuum distillation. **E. DINGEMANSE, J. FREUD, S. KOBER, E. LAQUEUR AND A. P. W. MCNICH** *Proc. Acad. Sci. Amsterdam* 33, 1206-9 (1930)—The method of fractionation is described and a schematic outline of the fractions obtained is presented. It is concluded that the male hormone can be brought to a high degree of purity through high-vacuum distn. At 90° with pressure of 0.001 mm. the hormone is found in the distillate, and at 80° in the residue. Fractional distn. offers the possibility for the sepn. of the male from the female hormone. Also in *Naturwissenschaften* 19, 166 (1931). **B. S. LEVINE**

The quantitative knowledge of cerebral lipoids. (A dissertation.) **ERICK BACKLIN** *Uppsala Läkareförening's Förhandl.* N. F. 35, 105-83 (1930)—The object was to det. the changes occurring in the cerebral lipoids in the psychoses. The chemistry of the lipoids of the brain and methods for the quant. study of cerebral lipoids are reviewed. A new method is described which requires only 30-50 mg. of the brain substance for the detn. of (a) cholesterol, (b) the fraction contg. unsatd. phosphatides in greater part, (c) the satd. phosphatides and (d and e) cerebrosides with satd. and unsatd. fatty acid radicals. Results obtained by this method indicated that it is reliable. The water content of the brain substance was detd. In the autolysis taking place in the cerebral

bipods during the first 24 hrs. postmortem, demonstrable changes actually occur

B. S. LEVINE

Effect of composition of air on the growth and mortality of the chick embryo. A. L. ROMANOFF. *J. Morphol. Physiol.* 50, 517 (1930); cf. *C. A.* 24, 4087.—Studies were made upon fertilized hen eggs. Continuous exposure to 0.4% of CO₂ in the air stimulated growth during the first part of the embryonic period. A high CO₂ and a diminished O₂ content decreased the size and increased the mortality of the embryos. Temporary exposures to high CO₂ concn. of 24- or 48-hr. periods decreased the size of the embryos with no apparent injuries.

C. M. McCAY

The phosphorus partition in normal whole blood. BERNHARD S. WALKER AND ANDREW E. HENTZINGER. *J. Lab. Clin. Med.* 16, 247-52 (1930).—The total P content of normal whole blood is 39.2 mg. per 100 cc. The inorganic phosphate P appears to be 9.7% of the total P, the acid-sol. P 63.1%, the acid sol. org. P (ester P) 58.3% and the lipid or lecithin P 36.0%. The P detns. were made by the method of Fiske and Subbarow (*C. A.* 20, 1092).

E. R. MAYN

The effect of splenectomy on changes in the blood at high altitudes. ANTONIO CHIATELLINO AND STEFANO GOLDBERGER. *Arch. sci. Biol. (Italy)* 15, 407-22 (1930).—The studies were made on 3 normal and 3 splenectomized dogs which had been kept on Colle d'Olen for 6 weeks. The percentage increase in red blood corpuscles for the normals was 5.0, 8.9 and 0.6, for the splenectomized animals 50.0, 22.3 and 53.3. The percentage increase in hemoglobin was, resp., 14.3, 19.5, 11.8 and 31.8, 59.0, 11.7. There was also an increase in leucocytes which appeared rather late, but there was no marked difference between the 2 groups. Blood smears from the splenectomized animals showed numerous Jolly bodies and a few normoblasts; the controls also presented a similar picture but with much less frequency. The resistance of the corpuscles was greater in the splenectomized than in the normal animals. On returning to lower altitudes, there was a marked diminution of all values which approached the initial values, the resistance of the corpuscles, however, remained higher than the initial value.

PETER MASTROCI

Gaseous exchange of the testicle. A. CHIATELLINO. *Arch. sci. Biol. (Italy)* 15, 433-44 (1930).—The respiratory exchange of the testicles of normal decerebrated, narcotized and curarized dogs was detd. In the first 2 groups, the blood flow in the organ was noticeably less than in the curarized group. In those narcotized the R. Q. was on an av. 45.7 cc. per min. per 100 g. of organ. The O₂ consumption was on an av. 3.47 cc. per 100 g. of organ, the CO₂ production was on an av. 3.00 cc. per min. per 100 g. of organ. The respiratory exchange in the left and right organs of the same animal corresponded perfectly. There were marked differences, however, between individual animals, in general with an increase in blood flow, there was a decrease in the respiratory exchange.

PETER MASTROCI

Variations in the nonprotein nitrogen of the blood during fatigue. UGO ROYCELLI AND S. CHIARELLA. *Minerva med.* 1930, 1, 65-70.—The purpose of this investigation was to det. the relation between muscular fatigue, hyperthermia and protein metabolism. From the results obtained on 6 subjects, the conclusion is drawn that the increase in nonprotein N of the blood which accompanies muscular fatigue (running) is independent of the energy spent and of the hyperthermia from fatigue. The increase in the nonprotein N is attributed to a change in equilibrium between the tissues and blood plasma.

PETER MASTROCI

Choline and adrenals. GASTENO VIALI AND TEODORO COMBER. *Arch. Biol.* 28, 25-32 (1930).—Injections of adrenaline cause the passage of choline from the adrenals into the circulation. The supradiaphragmatic section of the vagus has no influence on the choline content of the adrenals.

A. E. MEYER

Experimental determination of sex and the sterility by hormone action. B. PANIZZA. *Arch. Biol.* 28, 51-68 (1930).—See *C. A.* 25, 329.

A. E. MEYER

The metabolism of alcohol by man at high altitudes. A. BORNSTEIN AND A. LOEWY. *Biochem. Z.* 230, 51-67 (1931).—The alc. content of the blood rises more rapidly and to a higher level at high altitudes than in the plains, but after the max. concn. has been reached the loss of the alc. from the blood proceeds at the same rate at high or low altitudes. The mun. vol. as well as the frequency of the respiration is little affected under both conditions. The R. Q. was found to be regularly lower in the mountains than in the plains, and after the consumption of alc. it fell much more so than in the mountains the R. Q. often became that of pure alc. oxidation. Under neither condition did alc. exert any sp. dynamic action, except when sugar was given together with the alc.

S. MASTROCI

Phosphoribon, lactate acid formation and phosphatase activity in muscle pulp and

in muscle powder. J. BORNÁR AND BÉLA TANKÓ *Biochem. Z.* 230, 228-32(1931).—The lactic acid formation of muscle pulp decreases during preservation in direct ratio to its loss of ability to esterify H_2PO_4 . The H_2PO_4 -esterifying substance is completely retained in the powder prep'd. from muscle pulp with alc. ether, or with acetone ether. The inability of the muscle powder to form lactic acid is attributed to the diminished content of phosphatase, since the loss of phosphatase activity is parallel to the diminution in lactic acid formation.

S. MORGULIS

Fate of fructose in the animal organisms. II. Do the digestive juices cause a change of fructose to glucose? W. W. OPPEL. *Biochem. Z.* 230, 209-84(1931); cf. *C. A.* 25, 1858.—When a fructose sol'n is preserved under toluene in a thermostat with a mixt. of digestive juices (gastric, pancreatic and intestinal) at different pH values there is no formation of aldoses even after 20 hrs. Shorter preservation of similar mixts. but without toluene likewise fails to bring about any transformation of fructose to an aldose, although the intestinal sucrase causes considerable hydrolysis of sucrose in the same period of time.

S. MORGULIS

Studies on blood glucolysis. II. The pyrophosphate fraction and glucolysis. H. K. BARRENSCHEEN AND BÉLA VÁSÁRHELYI. *Biochem. Z.* 230, 330-46(1931), cf. *C. A.* 25, 1544.—The pyrophosphate content of different bloods shows considerable variation, increasing in the following order: horse < cow < dog < rabbit < guinea pig < man < pig. In man the pyrophosphate fraction constitutes relatively the highest portion of the acid-sol. org. P. Except for the blood of the pig, with its very large pyrophosphate content but no glucolytic power, the glucolytic ability of the various bloods seems to be correlated with their pyrophosphate content. During glucolysis there is an increase in the pyrophosphate which is entirely accounted for by the decrease in inorg. P, but when the glucolysis is accelerated by the addn. of SO_4 ions the pyrophosphate diminishes instead. Neither the feeding of large amts. of glucose nor the preliminary injection of insulin has any effect upon the pyrophosphate fraction of the blood P. The pyrophosphate is present only in the erythrocytes, apparently in some org. combination because it cannot be easily isolated. The blood and cells contain a pyrophosphatase which sets free P from pyrophosphate, the enzyme being inhibited in its action by F.

S. MORGULIS

The solubility, dissociation and tension of carbon dioxide in urine. FRITZ MAINZER AND MAROA BRUHN. *Biochem. Z.* 230, 395-410(1931).—Detns. made on 20 urines gave the following results: the soly. $\alpha(CO_2, 15^\circ) = 0.441 - 0.514$; $pK_1 = 5.81 - 6.30$; and $P(CO_2, 15^\circ) = 13.3 - 24.2$ mg. Hg. From this it seems that the CO_2 tension of the urine could not have the physiol. significance as representing the tension in the tissues.

S. MORGULIS

The cholesterol in structural combinations in protoplasm. V. Studies on adrenals from cattle. G. PFRIFFER. *Biochem. Z.* 230, 438-45(1931); cf. *C. A.* 24, 4828.—A special procedure is described for the hydrolysis of the adrenals with 2% NaOH. In 100 g. of fresh substance was found 0.2753 g. total cholesterol, of which 56.6% were cholesterol and oxysterol (the latter alone representing 17.5%) and 43.2% cholesterol esters. Cholesteryl palmitate was 2.3% and the oleate 9.1%; oxysterol esters were 31.9%. The high cholesterol content of the adrenals is thought to be due to the destruction of erythrocytes for the building up of the adrenal hormone. The endogenous synthesis of cholesterol in the adrenals is considered as improbable.

S. MORGULIS

The significance of the conjugated bile acids in fat resorption. IV. F. VERZAR AND A. VON KUTHY. *Biochem. Z.* 230, 451-7(1931); cf. *C. A.* 23, 4957.—The objection to the hypothesis that fat is absorbed in the intestinal tract as the water-sol. combination of fatty acids with bile acids could be based upon the fact that there is never enough bile acid present to convert all fatty acids into a diffusible state. Expts. made on isolated and washed out segment of intestines of living animals show that a given amt. of bile acid promotes the absorption of much more fatty acid than could be made diffusible in tests *in vitro*. It is therefore assumed that the bile acids become adsorbed on the epithelial cells of the mucous membrane where the fatty acids become converted into diffusible forms. This is corroborated by the finding that even hrs. after the administration of bile acids these can be demonstrated in the mucous membrane though not in the lumen of the intestine. The adsorbed bile acid-fatty acid combination is presumed to undergo hydrolysis, the bile acids being then ready to take up more fatty acids.

S. MORGULIS

The role of the pyloric sphincter in the control of gastric acidity. ROBERT ELMEN AND A. P. ROWLETTE. *Arch. Surg.* 22, 426-37(1931).

JOHN T. MYERS

The zinc content of the liver of the rat in relation to growth. GABRIEL BERTRAND

AND MME Y BRANDT BEALZEMONT *Compt rend* 191, 1410-1(1930); cf. C. A. 24, 5358.—With the same methods of technique and analysis as previously described, expts were made to det the Zn content of the livers of rats at various stages of growth at birth at 15 days (near weaning), at 30 days and at 7 months. It was shown that the quantity of Zn per 100 g. of liver decreased with age, there being about $3\frac{1}{2}$ times more of metal at the time of birth than in the adult rat. Because of this storage of Zn at birth, it seems that the role of Zn in nutrition is more important in the adult than in the new born. N. M. NAVLOR

The mucin of the articular fluids. CH ACHARD AND M. PIETTRE. *Compt rend* 191, 1412-4(1930).—In the expts described, synovial mucin is purified by repeated pptns from a water soln with acetone. It is first obtained as a thick, viscous mass but with further purification, as a dry powder. Synovial mucin averages 0.7% S, while the mean content of S in blood serum averages 1.4%. The sp viscosity of synovial mucin is 5.350, that of blood serum is 1.000. The synovial mucin described compares favorably with true mucin (1) by its soly in excess HCl and its slight soly. in excess AcOH and (2) by the fact that its content of sulfur is similar to that obtained for saliva mucus (Hammarsten) and for tendon mucus (Loelisch). N. M. NAVLOR

Threshold relationships of testis hormone indicators in mammals, the rat unit. CARL R. MOORE AND T. P. GALLAGHER. *J Pharmacol* 40, 341-45(1930); cf. C. A. 24, 3547.—The order of sensitivity in the rate of testis hormone indicators was (1) spermatozoon motility test, (2) prostate cytology test; (3) seminal vesicle, Cowper's gland and vas deferens tests, (4) elec. ejaculation test. The rat unit is defined as the minimal daily amt. required to maintain 50% of castrated animals in a normal condition, the indices being a normal prostate gland and seminal vesicle. One rat unit is approx equal to 6 bird units. C. RIEGEL

The chemistry of muscular contraction. OTTO MEYERHOF. *Lancet* 1930, II, 1415-22.—The hydrolysis of creatinephosphoric acid in muscle is of the utmost importance in muscle contraction. The lactic acid-forming system is analyzed and a "hexokinase" discussed. Similarities between the glycolytic function of muscle and the fermentative function of yeast are pointed out. The importance of phosphoric esters pyrophosphate and a deoxylic acid in muscle or enzyme expts is stressed. All vertebrates are found to contain creatinephosphoric acid in their striated muscles, and all classes of invertebrates utilize argininephosphoric acid. F. B. SEIBERT

The regulation of oxygen consumption in animals with variable alveolar gaseous tension. II. J. JORDAN. *Arch. néerland. physiol.* 15, 108-212(1930).—Diving birds, amphibia, molluscs and insects are considered. The O content of the lungs may vary more than 10% with only a slight difference in consumption. H. A. SOLTZ

Iron in normal and pathological tissues and its biological significance. I. Iron content of the organs of various classes of animals. KENJI KOTMA. *Nagoya J. Med. Sci.* 5, 34-48(1930).—In mammals and fishes, the Fe content of the spleen is always greater than that of the liver, but in birds, reptiles and amphibians the opposite is true. On the other hand, the glutathione content is greater in the liver of mammals and fishes than in the spleen and vice versa for birds, reptiles and amphibians. The amt. of both Fe and glutathione in the lungs of birds is relatively higher than in the other organs. II. Relationship of the iron content of the organs and the growth of animals. *Ibid.* 49-61.

The fluctuations in the Fe content of growing honeybees and silk worms parallel the cytochrome content, gradually becoming greater during the larva stage, decreasing during the pupa stage and increasing again in the imago stage. The amt. of Fe in growing chick embryos reaches its max. somewhat later than the glutathione and cytochrome content. The Fe content per g. dry substance in the organs of normal chickens was found to be: liver 0.1806, heart 0.1317, leg muscle 0.0870, brain 0.1444, kidney 0.0420, lung 0.0525 and blood 0.1468 mg. Egg yolk contained 0.0314 mg. Fe/g. dry substance, but no Fe was found in egg white. The water content of the organs of chick embryos decreased with growth. III. Effect of the time of year upon the iron content of animals that hibernate (toad) (*Bufo vulgaris japonicus*). *Ibid.* 62-70.—In general, the Fe content of the various organs of the toad is greatest in winter and least during the spawning time. The glycogen content parallels that of Fe, but the glutathione content is just the reverse. The liver always contains more than twice as much Fe as the spleen. A. WIEBEN

The carbon dioxide content of the blood in the new born. HOWARD F. KANE AND JOSEPH KREISELMAN. *Am. J. Obstet. & Gynec.* 20, 826-7(1930).—A study of 47 specimens indicates that the CO_2 of the blood in the new born is consistently high and increases with the degree of asphyxia. R. C. WILSON

Alkali reserve in blood during gestation. S. PASCALE. *Clinica Ostetrica* 32, 537-55

(1930); *J Am Med Assoc* 96, 76—From studies of 45 cases in various months of gestation, it is concluded that in the first 3 months of pregnancy the alkali reserve is lowered, but not markedly. The normal limits of alkali reserve are maintained between the fourth and eighth months. In the ninth month the reserve is about normal but tends to fall in the beginning of labor. R. C. WILLSON

Cholesterol and bilirubin contents of blood of mother and child. H. HINGLAIS AND J. GOVARTS. *Gynec et Obstet* 22, 137(1930). *J Am Med Assoc* 95, 1870—Hypercholesterolemia in the mother is frequent at term but it is by no means observed constantly. The cholesterol content of infants at birth is distinctly low. During the first few days after birth the blood cholesterol of the infant increases markedly. No relation was found to exist between the excess blood bilirubin at birth and the low cholesterol value. Bilirubinemia decreased in the first few days after birth. R. C. WILLSON

Organic and inorganic phosphorus content of blood in children. P. FEDOROV. *Vrachebnoe Delo* 13, 505-10(1930). *J Am Med Assoc* 95, 1789—The results given are from exams of 1000 samples of blood from 254 healthy children aged from 7 to 15 yrs. The av. inorg. P in blood of boys and girls was about the same, 3.97 mg per 100 cc. Org. P was 6.02 mg per 100 cc. No seasonal changes in org. P were observed. R. C. W.

Oxygen and carbon dioxide inhalations and their effect on circulation. C. KROETZ, Z. Kreislaufforsch 22, 611(1930). *J Am Med Assoc* 96, 80—In man O₂ inhalations do not increase the catabolic processes in the organism, there is, however, a high elimination of N₂. The latter reaches, or nearly so, the increase in the amt. of O₂ intake. In healthy subjects the arterial blood O₂ remains unchanged but the amt. of O₂ in venous blood is usually greatly increased. There is a definite increase in the min. vol. of the circulation. O₂ inhalations in cardiacs produce an increase in the O₂ tension of the tissue, a normalization of the O₂ content of the arterial blood and an elevation of the low min. vol. and beat vol. in decompensation. CO₂ inhalations increase the O₂ content of the arterial blood and increase the venous reflux and with it the min. vol. and the beat vol. R. C. WILLSON

G—PATHOLOGY

II GIBSON WELLS

Anaphylaxis with water-soluble specific substance from yeast-like fungi. T. J. KUROTSUKIN AND C. E. LIM. *Proc Soc Exptl Biol Med* 28, 223-5(1930)—The sensitizing power of the mycelial growth is distinctly higher than that of the budding cell growth of the same fungus. *Monilia penicillata* had a higher sensitizing capacity than *M. pilosus*, culturally, the former fungus differed from the latter in that its growth was accompanied by a rapid and rich formation of the mycelium. The guinea pig was used. C. V. BAILEY

Zinc and cancer. A. ZLATAROFF. *Bull soc chim biol* 12, 41(1930)—It is suggested that accumulation of Zn is one defense of the organism against the abnormal metabolism of the cancer cell, the Zn tending to inhibit the enhanced activity of proteolytic enzymes and catalase, to stimulate the decreased activity of peroxidase and to oppose the increased colloidal activity of the cancer cell, thus tending to reestablish normal cellular respiration. B. C. A.

Use of precipitated diphtheria toxin in the preparation of anti-diphtheria serum. P. SEDALLIAN AND MME CLAYEL. *Compt rend* 190, 1523-6(1930)—The ppt obtained when diphtheria toxin is adjusted to pH 4.7 (cf. *Compt rend soc biol* 97, 1543-5(1927), C. A. 22, 810, 23, 3500) is dissolved in a neutral peptone broth (Martin), and then injected (in increasing amts during 3 months) into horses. The serum obtained is of the same value (as an antigen) as that resulting from injection of the original toxin. B. C. A.

Potassium content of pleural discharges. DEMITRESCU-MANTE. *Compt rend soc. biol* 102, 971-2(1929)—Av. values are 25-35 mg per 100 cc. B. C. A.

Effect of various salts on the stability of diphtheria toxin. S. SCHMIDT. *Compt rend soc. biol* 103, 93-8, 98-100, 104-6(1930), cf. C. A. 24, 5855. B. C. A.

Role of the electrolyte in the diphtheria toxin-antitoxin reaction. Coagulation of the purified toxin in presence of various salts. S. SCHMIDT. *Compt rend soc. biol* 103, 101-3, 106-8(1930); cf. C. A. 24, 5855. B. C. A.

The influence of different salts on the stability of toxin and antitoxin. S. SCHMIDT. *Z. Immunitäts* 66, 506-18(1930), cf. preceding abstract—Diphtheria antitoxin is much more stable to the destructive action of salts than is toxin. The flocculating function of both substances is less stable than the antigenic or antitoxic properties.

Salts of acids which have strong oxidizing or reducing actions quickly destroy toxin and antitoxin. This is also true for salts of certain aromatic acids, as salicylic and cinnamic acids. JULIAN H. LEWIS

Occurrence of xanthine calculi in New Zealand sheep. T. H. EASTERFIELD AND J. A. BAUCS. *New Zealand J. Sci. Tech.* 11, 357-61(1930).—Calculi found in the kidneys of certain sheep consist of xanthine. The probable cause is a deficiency of minerals in the pasturage. B. C. A.

Hypochromic anemia with achlorhydria. THEODORE R. WAUGH. *Arch. Internal Med.* 47, 71-81(1931).—Mostly clinical. J. B. BROWN

The cholesterol partition of the blood plasma in parenchymatous diseases of the liver. EMANUEL Z. EPSTEIN. *Arch. Internal Med.* 47, 82-93(1931).—The cholesterol ester values of the blood plasma in acute parenchymatous damage of the liver are diminished or reduced to zero in proportion to the severity of the disease process. The cholesterol partition is normal in atrophic cirrhosis. J. B. BROWN

Chemistry and metabolism in experimental yellow fever in *Macacus rhesus* monkeys. III. Blood sugar and liver glycogen. A. MAURICE WAXMAN AND CLARA MORRELL. *Arch. Internal Med.* 47, 104-15(1931); cf. *C. A.* 24, 5835.—Monkeys with yellow fever showed hypoglycemia 24 hrs before death. Hypoglycemic convulsions during fasting could not be produced even with blood sugar as low as 45 mg %. In yellow fever, monkeys with hypoglycemia showed low liver glycogen. Adrenalectomy produced no hyperglycemic symptoms. J. B. BROWN

The potassium content of the hearts of persons dying from edematous and non-edematous conditions. L. C. SCOTT. *Arch. Internal Med.* 47, 116-21(1931).—There is no appreciable difference in the content of water, Na and K in the hearts of edematous and nonedematous persons. There is considerable individual variation. The Na content may be greater or less than K. J. B. BROWN

The blood lipids in nephrosis and chronic nephritis with edema. L. LICHTENSTEIN AND EMANUEL Z. EPSTEIN. *Arch. Internal Med.* 47, 122-7(1931).—Increased cholesterol, both total and ester, and phosphatides were found in the blood plasma of patients with various nephroses. The lipodermia was accompanied usually by diminution of serum proteins, especially albumin. Similar results were obtained on 7 cases of chronic glomerulonephritis. J. B. BROWN

Terminal hypoglycemia. B. G. SCHMIDT AND T. NELSON CARRY. *Arch. Internal Med.* 47, 128-34(1931).—Data of blood sugar in 33 nondiabetic patients dying in the hospital showed low terminal blood sugars (28-75 mg %) in 36% of the cases. Urea usually rose just prior to death. J. B. BROWN

The optical activity of cerebrospinal fluid in suppurative meningitis, and its lactic acid, sugar and chloride content. SYDNEY L. WRIGHT, JR., ELIZABETH F. HARR AND JOHN R. PAUL. *J. Clin. Investigation* 9, 443-61(1930).—The lactic acid content in the cerebrospinal fluid in suppurative meningitis is decidedly abnormal and usually related to the severity of the clinical symptoms. The ultrafiltrate of this fluid is usually *l*-rotatory. The lactic acid may be present as *l*-rotatory *d*-lactates. The ratio of concentration of Cl^- in the blood and spinal fluid is not altered. The sugar and lactic acid ratios are greatly altered, there being less sugar and more lactic acid. An extended bibliography is given. J. B. BROWN

The proportion of certain important inorganic constituents in the dying heart muscle. L. C. SCOTT. *J. Clin. Investigation* 9, 463-74(1930).—As a result of a large series of analyses of heart muscle from persons dying from a variety of diseases, there does not seem to be any relation between the water content, the content of ash, Ca, Mg, K and Na and disease. With the exception of Mg, the inorganic constituents are quite variable. Hearts from individuals dying from arteriosclerotic changes do not show high Ca content. A reversed ratio of K to Na may indicate an affected cardiac contractility. J. B. BROWN

The humoral origin of the paresis consecutive to the destruction of the suprarenals. J. E. ABELOUS AND H. LASSALLE. *Compt. rend. soc. biol.* 103, 475-7(1930).—The injection of blood from a frog paralyzed as a result of destruction of its suprarenals into a frog whose suprarenals have been extirpated results in paralysis and death. If injection is made into a normal frog, the same disturbances result after a very short delay. They are, however, transient, the poisons being gradually destroyed or neutralized by adrenaline. The considerable increase of muscular chronaxy in relation to that of the nerve implies a curarizing action of the poisons which accumulate in the organism when its suprarenal function is suppressed. B. C. BRUNSTATER

The protein equilibrium of blood serum in cerebral tumors. CH. ACHARD, M. BARIÉTY AND A. CODOUNIS. *Compt. rend. soc. biol.* 105, 346-7(1930).—A study of

10 cases of cerebral tumor showed that the protein equilibria of the blood sera were practically undisturbed. The range of values for total protein was 76.30–89.74 g. per 1000, for albumin, 39.10–58.84 and for globulin, 25.36–41.07 g. per 1000. The albumin-globulin ratio ranged from 0.95 to 2.12. B C BRUNSTETTER

Oxalaturia and increased oxalic acid secretion in urine in tuberculosis. JOSEF A LANGER AND TH LITTEG. *Beitr klin Tuberk* 72, 492–6(1929).—Among 150 tuberculous individuals examined there were 29.2% with a pos. oxalaturia, 48.2% with variable findings and 22.6% persistently neg. Of the latter there were about 61.7% with severe processes. Among children and youths under 16 years oxalate crystals were found less frequently. No effect of a febrile reaction was noted and a diet rich in oxalic acid was without effect. II J CORPER

Experimental studies on artificial immunization against tuberculosis. Hemolysin tests. P KALLOS AND E BAJZA. *Beitr klin Tuberk* 73, 323–4(1930).—In earlier expts. on artificial tuberculosis immunization in which tuberculous tissues were placed subcutaneously in sacs impervious to bacteria (*Beitr klin Tuberk* 71, 604(1929)) it was recommended that this method be used for immunizing purposes. In order to det. whether these sacs are suitable for the production of antibodies, sheep erythrocytes were placed in the sacs subcutaneously in rabbits and it was shown that the serum of these animals after a short time contained a fair titer of hemolysins. II J C

Bilirubin determination in the blood serum in chronic pulmonary tuberculosis. F WARNECKE. *Z Tuberk* 54, 321–7(1929).—As the result of an examn. of 134 patients with severe bilateral pulmonary tuberculosis there was found a diminution of bilirubin in the blood serum 0.002 mg. as compared to normal 0.005 mg. Toxic destruction of the liver parenchyma, of the reticulo-endothelium, insufficient supply of hemoglobin or O unsatn. of hemoglobin are looked upon as possible causes. The detns. were made by means of direct and indirect diazo reactions and with the Autenreith colorimeter. II, J CORPER

Blood sedimentation and white blood picture in the prognosis of pulmonary tuberculosis. HERMAN VOS. *Z Tuberk* 55, 431–7(1930).—Study of the blood picture and sedimentation revealed a disagreement between the 2 in 22% of 309 cases examd. In the cases where disagreement occurred the majority showed a parallelism of the clinical course and the leucocyte formula while the smaller proportion revealed an agreement of the sedimentation rate and the clinical course. Conclusion: The hemogram more rapidly reacts to the immuno-biologic condition of the body than the sedimentation rate. II, J CORPER

Studies on the tuberculin reaction. G PLATONOV AND S SHAYROVA. *Beitr klin Tuberk* 73, 435–41(1930).—In elaboration of observations made by others that certain substances possess the property of inhibiting or neutralizing the local tuberculin reaction, the authors used solns. of Na morrhuate and Na gynocardate, as well as solns. of Ca and Mn salts. After the amt. of tuberculin producing a reaction in the patient had been detd. these various reagents were added in different concns. Na morrhuate and Na gynocardate caused a definite diminution in the reaction while MnCl₂ increased it and CaCl₂ gave indefinite results. It is believed that the action of Na morrhuate and Na gynocardate is to be explained on the basis of the antitryptic properties of the unsatd. lipoids, while with CaCl₂ the possibility of the action through the nervous system cannot be excluded. On the basis of a parallelism between the local and general sensitiveness, it is reasoned that these expts. may point to a way of desensitizing lung tissue for therapeutic purposes. J H CORPER

Further observations on the serum calcium and plasma cholesterol in health and disease and on the blood chemistry in osteomalacia. T. A. HUGHES, D. L. SHIRIVASTAVA, P. N. SAHAI AND K. S. MALIK. *Indian J. Med Research* 18, 517–26(1930).—The serum Ca of residents in the Punjab is higher than that of normal inhabitants of temperate climates. The oral administration of vitamin D (Vigantol or Radiostol) or vitamin A (Radiostoleum) caused a rise of serum Ca in 4 cases of osteomalacia, but indefinite changes in serum P or plasma cholesterol. II EAGLE

Chemical composition of urinary calculi in rats. S RANGANATHAN. *Indian J. Med Research* 18, 609–613(1930).—The compn. of the urinary calculi of albino rats varies with the diet. Ca rich stones develop on a Ca rich diet, they contain it as CaCO₃, Ca(OH)₂ or both, are poor in P and N and contain no uric acid. CaCO₃ stones contain little H₂O, Ca(OH)₂ stones may contain as high as 49% H₂O. Mg-NH₄PO₄ stones develop on diets contg. no added lime. II, EAGLE

Chemical composition of the "nucleus" of urinary calculi. S RANGANATHAN. *Indian J. Med Research* 18, 613–7(1930).—The nuclei of vesical calculi are not homo-

geneous. They contain uric acid, urates and only traces of phosphates, their composition usually, but not necessarily, parallel the composition of the surrounding stone. H. E.

Renal diabetes and ketonuria. WILHELM BRAD. *Klin Wochschr* 9, 1621-3 (1930).—A description is given of an unusual case of renal diabetes, with development of ketonuria during pneumonia. H. EAGLE

Functional disturbances of the kidneys in diabetic coma. A. W. ELMER AND M. SCHUPP. *Klin Wochschr* 9, 1631(1931).—Contrary to the conclusions of Gottschalk and Muller, ketonemia and acidosis may play a very significant part in causing kidney insufficiency in diabetic coma. Ketonuria, decreased alkali reserve and ketonemia do not necessarily run parallel. A very severe ketonemia (> 300 mg %) is quite compatible with negligible ketonuria and a plasma CO_2 of 40%. H. EAGLE

Calcium gout and calcinosis universalis. HERMANN STEINITZ. *Klin Wochschr* 9, 1632-4(1930). H. EAGLE

The metabolism of progressive muscular dystrophy. GERTUO MOSBERG. *Klin Wochschr* 9, 2031-2(1930).—In 3 cases of infantile muscular atrophy which presented no clinical signs of endocrine disturbances there was (1) an increase in blood Ca, cholesterol and lactic acid, (2) abnormal blood glucose and NaCl curves following ingestion and (3) a normal insulin curve. H. EAGLE

Alimentary lipemia. A. HILSCH. *Klin Wochschr* 9, 2062-4(1930).—Alimentary lipemia is not observed in cases of portal obstruction (atrophic cirrhosis, ascites) or in disturbances of fat absorption of the most diverse kind (fever, cachexia, bile duct obstruction). Alimentary lipemia is therefore not a functional test of a cirrhotic liver, but of enteral fat resorption. Its absence is, however, one of the regular and early symptoms of atrophic cirrhosis. H. EAGLE

Specific skin reactions in patients with gonorrhea by means of specific protein-free substances derived from gonococci. WOLFGANG CASPER. *Klin Wochschr* 9, 2154-8(1930).—A protein free, type specific deriv of Type I gonococcus, carbohydrate in nature, gave specific skin reactions in gonorrheic subjects, apparently allergic in nature. Type II derivs gave no reaction with these patients, conversely, the Type I substance gave no reaction in a patient who was sensitive to Type II. H. EAGLE

Uremia in scarlatinal nephritis. A. ALTMANN. *Klin Wochschr* 9, 2252-4(1930).—The usual premonitory signs may be completely absent in cases of uremia following scarlatinal nephritis. The non protein N may be normal, blood pressure may not be high and there may be no oliguria, hematuria or albuminuria. A suggests the possibility that the convulsive uremia of scarlatinal nephritis may differ in essential respects from the quiet uremia of chronic nephritis. H. EAGLE

Hyperinsulinism with a hypoglycemic symptom complex. P. KRAUSE. *Klin Wochschr* 9, 2346-9(1930). H. EAGLE

Nephrosis and glucosuria. H. STRAUSS. *Klin Wochschr* 9, 2388-9(1930).—Previous communications have shown that there is a low threshold value for blood sugar in nephrosis following the peroral intake of glucose. The results of Fell, which indicate an abnormally high kidney threshold, were obtained by the intravenous injection of glucose. H. EAGLE

Percutaneous immunization with the Löwenstein diphtheria protective saline. HEINRICH BAAR AND HANS BENEDICT. *Klin Wochschr* 9, 2395-6(1930).—Eighty % of 106 children became Schick neg following percutaneous immunization. Inducing an erythema by means of a mustard plaster previous to the application of the toxoid saline increases the absorption. H. EAGLE

Experimental basis for the serological demonstration of a specific cancer antigen. H. LEHMANN IACUS. *Klin Wochschr* 9, 2430-3(1930).—A lipid which reacts specifically with a rabbit anti carcinoma serum can be obtained in the acetone insol. petr ether-sol fraction of the original carcinoma tissue. This fraction does not react with Wassermann pos sera and does not react with a group specific antiserum. H. EAGLE

Humoral changes in a case of trypanosomiasis (a contribution to the question of auto-antibodies). OPÖN FISCHER. *Klin Wochschr* 9, 2436-8(1930).—The serum of a patient with trypanosomiasis contained agglutinins at 0° for human blood cells of all groups including his own cells, and complement fixing antibodies against an extract of human red cells. F. interprets these as antibodies to the hosts' own red cells, due to changes induced by the infection. H. EAGLE

Observations on the sugar metabolism in a case of spontaneous levulosemia. P. A. HEERES AND HERMAN VOS. *Klin Wochschr* 9, 2440-1(1930). H. EAGLE

Influence of the benzene nucleus on the specificity of azoproteins. M. ADAMT. *Compt rend soc biol* 103, 539-40(1930), *Physiol Abstracts* 15, 260.—In the azoproteins

obtained by coupling diazotized aniline with serum protein, the protein fragment has lost its specificity. An active precipitin corresponding to the serum employed in production is quite free from activity toward the azoprotein. Replacement of the aniline mol by a benzidine or naphthylamine mol causes no modification of the specificity of the antibody. G. G.

Antigenic properties of gelatin. R. BRUYNOCHE AND P. VASSILIADIS. *Compt rend soc biol* 103, 543-4(1930), *Physiol Abstracts* 15, 260—Gelatin and gelatin aniline were injected into animals with the view of ascertaining whether they caused the production of deviating substances. Neither serum gave the slightest trace of deviation of alexin on gelatin coupled with aniline, but both fixed the alexin on gelatin by itself. The inactivity of the gelatin aniline serum toward the antigen cannot be attributed to a want of sensitizing substances, it is rather due to the nature of the antigen. G. G.

Antigenic properties of collagen and their modification under the influence of radium emanation. J. LOISELLEUR AND ACH. URBAIN. *Compt rend soc biol* 103, 776-8(1930), *Physiol Abstracts* 15, 258-9—The flocculation of collagen under the action of Ra is preceded by diminution of viscosity and lowering of the antigenic properties. G. G.

Experimental production of chicken sarcoma with arsenic trioxide. N. C. LACLAU AND C. PILLADO MATHEU. *Rev soc argentina biol* 5, 744-67(1929), *Physiol Abstracts* 15, 262—Carrel's results could not be repeated. Injections of As_2O_3 and embryonic pulp did not produce sarcoma. Cholesterol injections did not modify these results. No results were obtained by modifying the time of contact of arsenic and embryonic pulp, varying the pH and adding $ZnSO_4$ or ext. of rat sarcoma. Substitution of embryonic by spleen pulp also gave neg. results. Granuloma were produced by injecting clay. Subsequent injections of arsenic did not produce cancer. G. G.

Factors affecting the deposition of dental calculus. GEOFFREY H. SMITH. *Australian J. Exptl Biol Med Sci* 7, 45-77(1930)—Loss of CO_2 , changes in pH value and phosphate content of the saliva are not important factors in the deposition of dental calculi. A salivary phosphatase from desquamated epithelial cells which forms inorg. phosphate from org. esters is the chief factor. The optimum pH of the enzyme is 5.5. C. G. KNO

Rapid precipitation tests for syphilis and blood transfusion. B. S. LEVINE. *Am. J. Syphilis* 15, 81-9(1931)—Antigen pptn. is merely the first step of the complement-fixation procedure made visible. The recent pptn. procedures used in the lab. diagnosis of syphilis are relatively not greater as improvements over the early pptn. procedures than the cold incubation complement fixation procedures (Kolmer and similar tests) are over the early Wassermann tests. The improved pptn. procedures are not intrinsically specific for syphilis. They are empirical tests exactly as are the improved complement fixation tests. None of the antigen pptn. procedures is any more sensitive in reaction at any stage of the syphilitic disease than are the properly standardized cold incubation complement-fixation procedures. Antigen pptn. procedures, similar to those of complement fixation, yield pos. reactions, mostly of the weaker intensity, with serum of individuals affected with certain chronic low grade conditions other than syphilis. Antigen pptn. tests for syphilis have their zones of nonreactivity similar to the complement-fixation procedures. The colloidal conditions prevailing in some sera at times may favor complement fixation and not antigen pptn. and vice versa. The serum conditions which lead to nonreaction either with the pptn. or complement-fixation procedures or both are intermittent in the majority of instances. The proper precautions to be taken in testing blood sera before deciding upon their suitability for transfusion are given. B. S. LEVINE

Electrodialysis of antitoxin serums. FERNANDO MOOREN. *Rev inst bacteriol* 5, 695-702(1930)—An exptl. review and also original study in the electrodialysis of serum antitoxin with a view to its possible application on a large scale to the concn. of immune serums. B. S. LEVINE

Purification of antidiphtheric serum by thermal coagulation in the presence of salts. RAÚL WERNICKS AND FERNANDO MOOREN. *Rev inst bacteriol* 5, 713-7(1930)—Heating antidiphtheric serum for 10 min. at 70° in the presence of salts destroys its antitoxic properties. Such destruction increases with the duration of the heating and with the salt concn. The liquid fraction of the serum thus obtained generally is poorer in total protein in comparison with the proteins carrying the antitoxin than is the original serum. Therefore the ratio antitoxin units:1 g. protein is between 2 and 3 times as great in this liquid fraction as in the original serum, but it never exceeds 3 times its value. The yield in the antitoxin varies and generally reaches

60%, though it may be as high as 92-95%. More of the antitoxin is recovered in the presence of $(\text{NH}_4)_2\text{SO}_4$ and Na_2SO_4 than in the presence of NaCl . B. S. LEVINE

A contribution to the study of the precipitation of diphtheria toxin-antitoxin. FERNANDO MODERN AND RAEL WERNICKE. *Rev inst Bacteriol.* 5, 705-12 (T. A.) (1930)—Pptd toxin antitoxin requires a min. reaction of $\text{pH} = 3.6$ to dissolve completely. In acid medium the pptd. T. A. liberates the antitoxin, the optimum reaction being $\text{pH} = 3.2$ to 3.6. Temp. has an influence upon this reaction, the optimum being at 10-15°. The antitoxin unit content increases in the course of time from the instant the pptd. T. A. has gone into soln. The yield in the antitoxin units recovered from the T. A. in acid soln. increases with the diln. B. S. LEVINE

Plurality of the antigen contained in *B. anthracis*. A. SORDELLI AND V. DETLOFF. *Rev inst Bacteriol.* 5, 778-85 (1930)—With the aid of water or physiol. salt soln., preferably warm, it is possible to ext. from *B. anthracis*, cultivated on serum agar, a complex possessing complement fixing properties in the presence of that fraction of antitoxin serum which is sol. in H_2O and CO_2 . In this ext. there is also a substance (hydrocarbon) which possesses the properties of a precipitinogenic antigen. Both substances are relatively thermostable in neutral and alk. reaction. The complement fixing complex is destroyed by heat in acid medium. Both complexes are pptd. by alc. and acetone. The precipitinogenic substance is not thereby altered with regard to its antigenic properties, whereas the complement fixing fraction completely loses its antigenic properties. B. S. LEVINE

Studies in allergy. I. The specific activity of pollen oil. EDGAR L. MILFORD. *J. Allergy* 1, 331-3 (1930)—Oil extd. from dry ragweed pollen with the usual fat solvents produced skin reactions in only $\frac{1}{2}$ of the patients who reacted to aq. exts. of the pollen. Passively sensitized areas of the skin of normal people reacted to the aq. exts. when these sites had been exhausted by the oil ext. JULIAN H. LEWIS

Allergy considered as a special type of alkalosis. HARRY BECKMAN. *J. Allergy* 1, 496-500 (1930)—Because clinical conditions assoc. with acidosis either relieve or prevent allergy B. believes that a "potential alkalosis" is an important causative factor in allergy. He emphasizes the use of nitrohydrochloric acid in the treatment of hay fever. JULIAN H. LEWIS

Studies on pollen and pollen extracts. IV. The allergically active constituent in pollen oil. MAJORIE B. MOORE, HOBART W. CROMWELL AND EDMOND E. MOORE. *J. Allergy* 2, 6-10 (1931)—The non-dialyzable H_2O -free EtOH -sol. fractions of pollens have a skin reacting activity. From the results of focal passive sensitization and desensitization tests it is believed that the active constituent in pollen oil is identical with one or more of the antigens extd. by aq. solvents. Chem. tests of the active non-dialyzable constituent of pollen oil indicates that it may be glucosidic in nature. JULIAN H. LEWIS

Further studies on the precipitation test for syphilis. EMIL WATTS. *J. Infectious Diseases* 47, 355-8 (1930)—The comparative study of Kahn's antigens with regard to the temp., time and ratio of tissue to solvent employed in their prepn. led to the prepn. of an improved antigen. Dry beef heart is extd. with EtOH in the ratio of 1:13 for 40 min. in the icebox. The ext. is discarded and the tissue dried. To the dried tissue 95% alc. is added in the ratio of 1:5. The alc. extn. is carried out at 37° for 3 days. This ext. is then filtered and left overnight at room temp. The ext. is refiltered and 0.6% cholesterol is added. This antigen is simpler to prep., has a wider sp. zone and does not require addnl. sensitizing or correcting. JULIAN H. LEWIS

Serologic studies on the proteins found in casein. D. C. CARPENTER AND G. J. HUCKER. *J. Infectious Diseases* 47, 435-42 (1930)—By extn. with acid 70% alc. and by fractionation with K. oxalate 3 proteins having mol. wts., resp., of 98,000, 183,000 and 375,000 were sep'd. from crude casein. These proteins were clearly distinguishable from one another by serologic reactions. It was also found that the alc.-sol. protein of Osborne and Wakeman is not the same as the acid alc.-sol. protein having a mol. wt. of 375,000. JULIAN H. LEWIS

Precipitation with cerebrospinal fluids. A. M. MALLOY, R. L. KAHN AND LUCY WESTALL. *J. Infectious Diseases* 48, 203-11 (1931)—As is true with the Kahn reaction, the colloidal Au and mastic reactions depend on the globulin fraction of the spinal fluid, the albumin fraction giving neg. results. The globulin fraction gives results somewhat more sensitive than the unfractionated fluid, undoubtedly because of the removal of the protective action of the albumin. After a Kahn test has been performed with a soln. of globulin obtained from the spinal fluid and after removal of the ppt. by centrifugation, the supernatant soln. continues to give colloidal Au and mastic reactions, indicating that the pptn. substances in the globulin soln. (or the

conditions of this soln.) responsible for the Kahn reaction are not the same as those responsible for the colloidal Au and mastic reactions. Heating the globulin fraction at 60° up to 2 hrs. exerts little effect on the 3 reactions. When heating is prolonged to 24 hrs., there is a tendency toward reduction in sensitivity, especially in that of the Kahn reaction. Heating the globulin exerts comparatively little reduction in the sensitiveness of the colloidal Au and mastic reactions and considerable reduction in that of the Kahn reaction. After heating at 70° for 30 min., there is reduction in the sensitiveness of the 1st 2 reactions and marked reduction in that of the Kahn reaction. These heating expts. also indicate that the conditions governing the Kahn reaction are not identical with those governing the colloidal Au and mastic reactions. J. H. L.

Detoxifying and disinfecting properties of sodium salicylate. Its action on diphtheria and tetanus toxins and on streptococcal "toxic" filtrates. KONRAD BIRKHAUG. *J. Infectious Diseases* 49, 212-25 (1931).—A satd. soln. of Na salicylate has the power of neutralizing diphtheria and tetanus toxins *in vitro* without destroying their antitoxinogenic capacity. Neutralization proceeds in direct proportion to the concn. of Na salicylate, the duration of contact and the degree of temp. Simultaneous injection of pure toxin and the required amt. of Na salicylate that produces a neutral mixt. *in vitro* fails to avert death from diphtherial toxemia, although the period of survival is slightly prolonged. No neutralizing effect is exercised by the satd. soln. of Na salicylate on the toxic filtrates of hemolytic or nonhemolytic streptococci, a feature that seps. these "toxins" from the true bacterial exotoxins of *C. diphtheriae* and *C. tetani*. The bactericidal action of Na salicylate on commonly encountered pathogenic microorganisms is about 1/10 that of phenol. The precise mode of action of Na salicylate *in vivo* remains obscure. Its dual capacity to exercise antitoxic and antiseptic action *in vitro* suggests the hypothetical possibility that its therapeutic success in certain infectious diseases is effected by weakening the pathogenic microorganisms and their capacity to secrete the deleterious exotoxins. JULIAN H. LEWIS

Titration of the antigenic properties of diphtheria toxin *in vitro*. E. HOEN AND L. TSCHERTKOW. *Z. Immunitäts* 66, 359-93 (1930).—Flocculation tests are not adapted in the titration of antitoxin binding values of diphtheria toxin and its derivs. but the ring pptn. test gives const. and accurate results. JULIAN H. LEWIS

The action of phenol-alcohol on serological reactions. G. SOLLAZZO. *Z. Immunitäts* 66, 424-42 (1930).—Phenol and alc. added to lipid antigens (sp. lipoids, Wassermann and Forssman antigens) make them more reactive in immune reactions. Phenol alone is more active than alc., but a mixt. of the 2 is most active. J. H. L.

The identity of heterogenetic antigens. K. A. FRIDDE AND L. A. SCHWARZMAN. *Z. Immunitäts* 66, 454-64 (1930).—Although 9 different heterogenetic antigens have been described the authors believe they are all the same antigen. Proof of this opinion is given in adsorption expts. JULIAN H. LEWIS

The presence of Forssman heterogenetic antigen in the blood corpuscles of the guinea pig. G. SOLLAZZO. *Z. Immunitäts* 66, 465-84 (1930).—Heterogenetic antiserum and antiserum obtained by immunizing with sheep blood react specifically with alc. exts. of guinea pig red blood cells. The hemolytic amboceptor for guinea pig red cells in these sera is difficult to activate with guinea pig serum but can be activated readily with rabbit serum. Guinea pig red cells adsorb antibodies from heterogenetic antisera slowly. Repeated exposures of red cells to the antiserum must be used. These expts. prove that guinea red cells contain Forssman antigen. JULIAN H. LEWIS

The antagonistic influence of glucose in anaphylactic shock. K. M. DVOLAITSEK-BARISHOVA AND M. S. GOLDBERG. *Z. Immunitäts* 66, 485-90 (1930).—By perfusion expts. it is found that glucose in 5% concn. has the property of inhibiting the sp. action of antigen on the vessels of a sensitized guinea pig. This action of glucose is ascribed to the fact that glucose has the ability to prevent the change in the grade of dispersity of the colloids in muscle cells that result from the combination of antigen with antibody within the territory of the cell. JULIAN H. LEWIS

The occurrence of the Thomsen antigen in human organs. M. N. SINYUSHINA. *Z. Immunitäts* 66, 491-5 (1930).—The receptor which is developed by human erythrocytes exposed to the growth product of certain bacteria and which makes them agglutinable by all sp. agglutinating sera is not developed under the same conditions by cells of human organs. JULIAN H. LEWIS

The appearance of organ specific brain substances in ontogenesis. E. WITEBSKY AND G. SOLLAZZO. *Z. Immunitäts* 67, 1-8 (1930).—The brains of guinea pig embryos 4-6 cm. in length and of human fetuses of the 3rd to 4th months react with sp. anti brain sera. In earlier stages the brains do not react. The sp. antigen of embryonal brains is heat stable and alc. sol., which indicates it is lipoidal in nature. J. H. L.

Potassium and calcium in anaphylaxis. M A KUSCHINAEV. *Z. Immunität* 67, 9-14(1930)—After anaphylactic shock there is a marked increase in the K and a decrease in the Ca contents of guinea pig serum. Anaphylactic shock is inhibited in sensitized guinea pigs if CaCl_2 is injected along with the antigen. Guinea pigs injected with isogenous or autogenous serum mixed with KCl or KI show symptoms of anaphylactic shock. Anaphylactic shock is undoubtedly concerned with disturbances of the Ca K equil in the vegetative nervous system. JULIAN H. LEWIS

The alkalosis of vomiting and the acidosis of advanced renal disease. T. G. H. DRAKE, P. MARSH AND J. L. GAMBLE. *Am. J. Diseases Children* 40, 705-17(1930)—The chloride ions lost by vomiting can be restored by the intravenous administration of physiol. NaCl soln and the increased concn of bicarbonate ion thus reduced to the normal value. The intravenous administration of 10% dextrose soln aids in restoring the renal function impaired by the progressive dehydration brought on by vomiting. The lowered concn of bicarbonate ion assocd with nephritis can be corrected by the intravenous injection of NaHCO_3 unless the impairment of renal control has already become too far advanced. Charts are included which give the quantitative compn of the plasma in the types of acidosis and alkalosis discussed and which illustrate quantitatively the changes in compn produced by the injection of H_2O and other substances. E. R. MAIN

Cerebrospinal fluid of premature infants. Results of a series of spinal punctures in one hundred and seventy cases, with special reference to the origin of physiologic xanthochromia of the new-born infant. JEROME GLASER. *Am. J. Diseases Children* 40, 741-52(1930)—Lumbar puncture during acute meningeal congestion usually yields cerebrospinal fluids which are definitely hazy but which do not contain a sufficiently large no. of erythrocytes to show a pink coloration. The yellow spinal fluids give a pos. van den Bergh reaction in 55% of the fluids examd. A pos. benzidine test on the supernatant fluid obtained by centrifugation of bloody fluids always indicates meningeal hemorrhage, but a neg. test does not rule out hemorrhage. The sterus index of both the blood serum and the spinal fluid is highest during the 2nd and 3rd weeks of life. E. R. MAIN

Bile-salt hemolysis in new-born infants and its inhibition by the blood serum. HEYWORTH N. SANFORD, MARIAN M. CRANE AND FLEANNOR I. LESLIE. *Am. J. Diseases Children* 40, 1039-44(1930)—The erythrocytes of normal, new born infants are more resistant to hemolysis by bile salts than those of adults and the serum gives greater protection against hemolysis. There appears to be no variation from the normal in idiopathic icterus, pathologic jaundice due to congenital syphilis or in hemorrhagic disease of the new born. E. R. MAIN

Pentosuria in children, with laboratory data on four cases. ALFRED E. FISCHER AND MIRIAM REINER. *Am. J. Diseases Children* 40, 1193-1207(1930)—Pentosuria in children is not assocd with an increased amt. of nonfermentable reducing substances in the blood or with changes in tolerance for glucose. The blood pentose curve of pentosuric children following the ingestion of aylose is similar to that of normal children. A small quantity of the ingested pentose may be accounted for in the urine and feces. The fate of the remainder is undetd. The excretion of ingested pentose is slightly greater and more prolonged than that characteristic of normal children. Pentosuria appears to be the result of a metabolic defect and not of an increased permeability of the kidney for pentoses. E. R. MAIN

The gastric and duodenal contents of normal infants and children. The duodenal enzyme activity and the gastric and duodenal reactions [hydrogenation]. THEODORE G. KLUMPP AND A. VICTOR NEALE. *Am. J. Diseases Children* 40, 1215-29(1930)—The pH of the gastric contents of children under 2 yrs. is approx. 4.1 and decreases to 2.8 in children of 9-12 yrs. The presence of true achlorhydria appears to be uncertain unless the histamine test is employed, since many of the infants and children show no free HCl during fasting and a smaller no. show none after a test meal of cream. The pH of the duodenal contents during fasting tends to remain at 7.0 and is decreased to approx. 5.0 after the test meal. The amylolytic enzyme activity of the duodenal contents during the 1st yr. of life is very low and then increases abruptly, exceeding the av. values for adults. The lipolytic activity is consistently low throughout infancy and childhood. The proteolytic activity is high. The gastric acidity and the activity of the lipolytic and amylolytic enzymes are slightly increased after the ingestion of the test meal of cream. E. R. MAIN

Inorganic blood chemistry in the osteochondritides. JOSEPH BUCHMAN AND ISAAC F. GITTLEMAN. *Am. J. Diseases Children* 40, 1250-6(1930)—The Ca, Na, K,

Mg and inorg. P contents of the blood serum in osteochondritides are within the normal limits. Rickets does not appear to be a causative factor. E. R. MAIN

Creatinuria in tuberculosis. I. Creatinuria in tuberculous men. JAMES T. THOMPSON *J Lab Clin Med* 16, 5-11(1930)—Creatinuria often appears to be assocd. with pulmonary tuberculosis. E. R. MAIN

Phosphorus metabolism. II. The distribution of phosphorus in normal and cancer bloods. GUY E. YOUNGBURG AND MAMIE V. YOUNGBURG *J Lab Clin Med.* 16, 253-9(1930), cf *C. A.* 25, 1857—The blood in cancer appears to vary from normal blood in the following respects: The ester P content of the corpuscles is increased 20.5%, the lipid P is increased 17.2% and the inorg. phosphate P is decreased 15.6%. It is not yet detd. whether the changes are characteristic of cancer only. E. R. MAIN

The leucocytic reaction to injections of a nonspecific protein. RUDEN L. LARSEN AND A. A. JANSON *J Lab Clin Med* 16, 362-9(1931)—The leucocyte content of the blood appears to be progressively increased in normal individuals following 6 injections of a non sp. protein prepn (Proteolac). A slight, temporary increase may follow one injection. The increases are not accompanied by local reactions, rise in temp. or chills. The size of the dose may not be a determinative factor. E. R. MAIN

A microscopic slide precipitation test for syphilis with spinal fluid. B. S. KLINE AND C. R. REIN *J Lab Clin Med* 16, 399-404(1931)—A microscopic slide pptn test is described in which spinal fluid is used directly with an antigen consisting of a cholesterolized alc. ext. of beef muscle. The test is somewhat more sensitive than the blood Wassermann test with the same antigen. Only fluids which reduce Benedict's soln. may be used. E. R. MAIN

The glucosuria of hyperthyroidism and its clinical significance. I. M. RABINOWITZ. *Ann Internal Med* 4, 881-96(1931). JOHN T. MYERS

Extra-insular (central) glucosuria with hyperglucemia following epidemic encephalitis. I. W. HELD, A. ALLEN GOLDBLOOM AND JULIUS CHASNOFF. *Ann. Internal Med* 4, 897-908(1931). JOHN T. MYERS

The behavior of some biochemical components of the blood and the spinal fluid in eclamptic uremia. L. CANNAYO *Biochim therap. spec.* 17, 478-85(1930)—Observations of 2 cases of eclampsia are reported. The NaCl in the blood is normal; it is increased in the spinal fluid. The hyperglucemia observed is supposed to be connected with the muscular convulsions. A. E. MEYER

Studies on experimental protein hemotoxic anemias. AAGE NYFELDT. *Folia Haematol.* 42, 129-59(1930). JOHN T. MYERS

Acidosis and hyperthyroidism; variations of the alkali reserve in the red corpuscles. ENRIQUE CANTILLO *Presse med.* 38, 722-3(1930)—Thyroid hyperfunction causes a decrease of the alkali reserve in the red corpuscles to the limit of 50%. In hypothyroidism no change was observed. The increased metabolism in the former case causes an over production of acids. A. E. MEYER

Cataract and calcium metabolism. R. RIVOIRE *Presse med.* 38, 723-7(1930).—The coincidence of cataract and diseases caused by Ca deficiency is emphasized. A. E. MEYER

Protein and protein quotient in the serum of tuberculous patients. ANDRÉ DUFOUR, ROBERT AND MOREAU *Presse med.* 38, 843-5(1930)—For the investigation of the proteins in serum, proceed as follows: Sep. serum from the clotted blood and dil. with 9 parts H₂O contg. a trace of AcOH. Within 24 hrs. the nucleoproteins are pptd. Filter and add to 20 cc. of the filtrate 5 g. NaCl, boil, add 6 to 8 drops of Cl₂CCO₂H (20%), cool and centrifugalize. Wash the ppt. several times with H₂O, at last with EtOH, collect on a filter, dry at 100° to 105° and weigh as total proteins. Neutralize 40 cc. of the dil. serum and ppt. the globulins with MgSO₄. Fill up to a vol. with a concd. soln. of MgSO₄, filter and use a quantity of the filtrate corresponding to 3 cc. of serum. Ppt. by heat and proceed as previously. This gives the serines, the globulins are calcd. by difference. The quotient serin % globulin % is 1.6 to 2.0 in normal persons. A high quotient gives a favorable prognosis in tuberculosis, a low one is unfavorable. But other factors have to be considered also to obtain a judgment of value. A. E. MEYER

The different forms of flocculation of colloids. AUGUSTE LUMIÈRE *Presse med.* 38, 873-5(1930), cf *C. A.* 24, 5570—The general characteristics of colloids are described. Disease is considered as a disturbance of the colloidal system. A. E. MEYER

Psychosis and blood lipoids. Quantitative variations of total cholesterol and total fatty acids in the blood. I. Manic-depressive psychosis. SVEN STENBERG *Acta*

med scand 71, 558-57(1929); cf. *C. A.* 24, 651.—For normal men and women the total cholesterol and total fatty acids were 0.1444 and 0.1524%, and 0.2466 and 0.3050%, resp. Both in the manic and melancholic phases of the manic-depressive psychosis there are notable changes in total cholesterol and total fatty acid of the blood, the values falling back to normal after recovery. S. MORGULIS

The brown pigment of hemochromatosis. MAATIN JACOB *Biochem Z* 230, 225-7(1931).—The liver obtained at autopsy of a case of hemochromatosis was finely cut to ext. the pigment. This was insol in alc, ether, $(\text{CH}_3)_2\text{CO}$ or CHCl_3 . Glacial acetic acid became strongly colored and the soln. was difficult to filter, but if the mixt was boiled a colorless filtrate was obtained. Neither did the pigment pass into benzene, phenol, toluene or CCl_4 . The pigment is very sol in 0.85% NaCl. If the liver pulp previously treated with NaCl soln is extd. with H_2O , it is possible to ppt. a brown substance by complete satn with $(\text{NH}_4)_2\text{SO}_4$. This was washed with satd. $(\text{NH}_4)_2\text{SO}_4$ suspended in H_2O and dialyzed until the Nessler reaction became neg. The substance is then dissolved by the addn. of NaCl and on filtration a clear brown soln. is obtained. Of the 691 mg dry residue in 100 cc. soln. 154 mg was ash; 91 mg N and 66 mg Fe. The relation between N and Fe is thus 4:1. S. MORGULIS

Urinary creatinine as a test of renal efficiency. A. Muta CRAWFORD *Lancet* 1930, II, 1177-79.—In a great variety of cases of nephritis, which are tested and described, the creatinine excretion for a 24 hr. period was detd. and would seem to be a reliable prognostic guide. F. B. SEIBER

Nature of the chill-producing principle in antipneumococcus serum. ALBERT B. SABIN AND G. B. WALLACE *J. Exptl Med* 53, 339-62(1931).—The chill principle appears to be formed only in blood which has been allowed to stand. The expts. reported do not show that the reaction is dependent on formed elements, fibrinogen or lipoids. Anticoagulants, filtration, dialysis and moderate heating are without effect in removing the principle from the soln. contg. it. By changes in the NaCl and H ion concn. in antibody solns. an acid globulin and an alk. globulin fraction may be obtained. The acid globulin fraction, whether or not phospholipin is present, contains the greater part of the chill principle and a small part of the antibody substance, the alk. globulin fraction contains the greater part of the antibody substance and a smaller part of the chill principle. The acid globulin fraction is not itself the chill principle but serves as a carrier of this, probably through an adsorptive process. C. J. WEST

Immunization with mixtures of poliomyelitis virus and aluminum hydroxide. C. P. RHODES *J. Exptl Med* 53, 399-401(1931).—See *C. A.* 25, 738. C. J. W.

Iron and copper in various tissues in acute myeloid leucemia. KENZU KOJIMA AND SHINZO KOSAKA *Nagoya J Med Sci* 5, 71-4(1930).—K. and K. tabulate results obtained in analyzing liver, spleen, heart muscle, kidney, cerebrum and blood in myeloid leucemia. The spleen was richest in Fe and the cortex of the kidney and brain contained more Fe than the marrow. However, the liver was richest in Cu and the cerebrum cortex contained nearly 3 times as much Cu as the marrow. A. W.

Normal urine sugar in cystoscopic examinations. RUSSELL RICHARDSON *Am J Med Sci* 180, 171-7(1930).—Normal urine sugars in specimens obtained by ureteral catheterization usually show approx. equal amts. of sugar from the 2 kidneys in the absence of any disease which damages kidney efficiency. The sugar from a badly damaged kidney is very much decreased as compared with that from a normal kidney. R. C. WILLSON

The hydrogen-ion concentration of the blood in eclampsia. H. J. STANDER AND N. J. EASTMAN *Am J Obstet & Gynec* 20, 822-5(1930), cf. *C. A.* 24, 884.—Immediately following an eclamptic convulsion the blood pH usually increases, often to a true acidosis. R. C. WILLSON

Studies in epilepsy. XI. The calcium content of the blood and of the spinal fluid. WM. G. LENNOX AND MARGARET B. ALLEN *Arch Neurol Psychiatry* 24, 1199-1203(1930), cf. *C. A.* 24, 5372.—Detns. were made in 77 epileptics. Av. concns. and av. spinal fluid serum ratios were within normal limits, although there was an abnormally wide distribution of values, with a tendency toward low spinal fluid measurements. In $\frac{1}{2}$ of the patients the ratio of spinal fluid to serum Ca was less than 45%. R. C. WILLSON

Acetone bodies in normal pregnancy and in the toxemias of pregnancy. H. J. STANDER AND J. F. CANOEN *Bull Johns Hopkins Hosp.* 47, 382-90(1930).—The total acetone bodies (free acetone, acetoacetic acid and β hydroxybutyric acid) in the blood of normal nonpregnant women varied between 11 and 18 mg. acetone per l. blood after correction for lactic acid, the av. being 15.8 mg. In normal pregnant women

corrected acetone bodies varied between 11 and 21 mg., with an av. of 16.6 mg. Corrected values in pregnancy with nephritis were 8.6 to 48.8 mg., and in eclampsia they averaged 50.2 mg. R. C. WILLSON

Alkalinity of the blood in eczema. B. S. BIKKENIN AND N. N. YASNITSKY. *Kazansky Meditsinsky Jurnal* 26, 703-8(1930); *J Am Med Assoc* 95, 1516.—Urine pH was normal but that of the blood was found to be from 0.1 to 0.5 above the av. R. C. WILLSON

Acid-base equilibrium in diseases of the heart and kidneys. I. M. LIPETS AND M. M. PISMAREV. *Kazansky Meditsinsky Jurnal* 26, 778-86(1930); *J Am Med Assoc* 95, 1629.—A decrease in the alkali reserve and a parallel decrease in the alveolar CO_2 were observed during the decompensation period in both heart and kidney diseases, usually more pronounced in the latter and in proportion to the severity of the condition. The Cl content of the urine remained unchanged. The blood urea was increased in almost all patients with kidney diseases, but in only a few patients with cardiac complaints. R. C. WILLSON

Peculiarities of glucemic reaction in early infancy. V. L. STYRIKOVICH. *Khirurgiya Materinstva i Detsstva* 1, 28-42(1929); *J Am Med Assoc* 95, 1780(1930).—S studied the various phases of the glucemic reaction and recognized the following main types and subtypes, according to the curves they give: (1) with a single rise, encountered in only 37% of cases, (a) to a considerable height with a more or less rapid descent to normal, (b) to a moderate height, remaining there for a time (1 hr. or longer), with sometimes a wave-like fluctuation, and thereupon dropping—in some cases with a tardy hypoglycemia, (c) to a modest height with subsequent drop to a lower level on an empty stomach; (2) with 2 rises (a) with the second rise lower than the first, and the second wave of smaller vol than the first, (b) with the second rise attaining or surpassing the level of the first, and the second wave equal to or exceeding the first, (c) with the second rise rapid, returning to normal after a short time; (3) with 3 rises—a rare form. R. C. WILLSON

Absorption of calcium and phosphorus in experimental rickets. F. PEOLA AND G. GUASSARDO. *Riv. clin. pediatrica* 28, 633(1930); *J Am Med Assoc* 95, 1541.—Normally the absorption of Ca is from 50 to 60% of the amt. taken in and that of P from 70 to 80%. In exptl. rickets that of Ca remains normal while that of P falls to 60%. This decrease appears later in the disease, however. R. C. WILLSON

Benzene derivatives of intestinal putrefaction in the etiology of cancer. E. GRANGEJEAN. *Schweiz. med. Wochschr.* 60, 823-6(1930); *J Am Med Assoc* 95, 1540.—Benzene derivs of intestinal putrefaction resulting from diets rich in animal proteins exert a continuous growth stimulation on the healthy cells, and gradually the constantly irritated tissues develop into malignant tumors. Benzene derivs are also formed wherever in suppurative processes disintegration of protein by anaerobic bacteria takes place, as in the lungs, female genitalia, abscesses of the teeth and ulcers. The theory that benzene derivs are significant in the etiology of cancer is supported by the fact that occupational cancers are most frequent in workers who come into contact with benzene and coal-tar products. R. C. WILLSON

Cardiac insufficiency; chemical changes in the blood. E. BRECHER. *Z. Kreislaufforsch.* 22, 638-40(1930); *J Am Med Assoc* 96, 80.—The amino-N content of the blood in cardiac insufficiency is almost constant and varies only slightly from the normal. Occasionally the creatinine content is increased. The blood serum does not contain an increased amt. of urochrome or urochromogen. The uric acid, urea and residual N are frequently increased but not so much as in renal insufficiency. There is no increase in the aromatic substances. Phenol, cresol and its derivs, and indican are not increased. R. C. WILLSON

H—PHARMACOLOGY

A. N. RICHARDS

Effect of some imidazoles on gastric secretion. J. P. BURGERS AND A. C. IVY. *Proc. Soc. Exptl. Biol. Med.* 28, 115-6(1930).—Imidazole, imidazolepropionic acid, imidazolealdehyde, 1-imidazolelactic acid and imidazoleacrylic acid in doses of 1 mg to 5 mg have no effect on gastric secretion or on blood pressure in the dog. C. V. BAILEY

Effect of calcium chloride injections on blood sugar of normal and jaundiced dogs. ALLAN G. REWBIDGE AND EDMUND ANDREWS. *Proc. Soc. Exptl. Biol. Med.* 28, 126-7(1930).—The intravenous injection of a 10% soln. of CaCl_2 caused a rise in blood sugar nearly as marked as that following the injection of sugar itself and tended to last

longer. Sugar and Ca seem to have a reciprocal influence on each other and one cannot say which is of greater importance in the control of blood coagulability. C. V. B.

Rate of urinary arsenic excretion after giving acetarsone ("stovarsol") and "carbarsone" by mouth. MET-YU CHU, H. H. ANDERSON AND C. D. LEAKE. *Proc Soc Exptl Biol Med* 28, 145-8(1930)—Acetarsone (3 acetyl-amino-4 hydroxyphenyl-arsonic acid) and carbarsone (4-carbamoylphenylarsonic acid) are slowly excreted in the urine after oral administration. A normal man gave 0.5 g. of acetarsone excreted 20% of the ingested As in 72 hrs., given a similar dose of carbarsone he excreted 8% in 52 hrs. Another subject excreted 7% of the As in 24 hrs. following 0.5 g. of acetarsone and 13% in 42 hrs. following the ingestion of 0.5 g. of carbarsone. C. V. B.

The pharmacology of inflammation. I. Technic. LOUIS HIRSCHMAN AND A. G. MULINOS. *Proc Soc Exptl Biol Med* 28, 168-70(1930)—One drop of a 15% soln. of essential oil of mustard in mineral oil, instilled into the conjunctival sac of the rabbit, induces blepharospasm for 1 hr., lacrimation for 0.25 hr., immediate injection of the conjunctival blood vessels lasting for 48 hrs., pin point miosis which passes off in 1 hr. and edema which is max. in 0.5 hr. Complete recovery usually occurs in 7-14 days. The method lends itself to studies of the inhibitory and corative effects of drugs on inflammatory processes. C. V. BAILEY

Further investigations concerning the stimulating effect of anterior pituitary gland preparations on the thyroid gland. LAO LOER, R. B. BASSETT AND HILDA FRIEDMAN. *Proc Soc Exptl Biol Med* 28, 209-13(1930)—Guinea pigs were injected daily with 1 cc. of an acid ext. of anterior pituitary gland for 5 consecutive days. On the 6th day the thyroid gland was markedly hypertrophied, the cells were mitotic and the colloid had largely disappeared. Retrogression progressed regularly and the gland had a normal appearance at the end of 31 days. The administration of thyroid substance diminished but did not prevent this hypertrophy. The injection of anterior pituitary ext. is more potent in producing hypertrophy of the thyroid gland than is the removal of a great part of the gland or the administration of stimulating doses of KI. C. V. BAILEY

Comparison of physiological action of aspidospermine and quebrachine. RAYMOND-HAMET. *Compt rend* 191, 157-9(1930)—The hydrochlorides of these alkaloids isolated from the bark of *Aspidosperma quebracho blanco*, contrary to reported expts. act differently on the sympathetic or vagus nervous systems. B. C. A.

Secretion of dextrose under the influence of phosphates and sulfates. J. VAN DE VELDE. *Compt rend soc biol* 102, 506-9(1929)—Injection of a soln. of Na_2HPO_4 into a dog after injection of dextrose accelerates the secretion of sugar in the urine, the blood corpuscles giving up more sugar than the plasma. Na_2HPO_4 has practically no effect on the pH of the blood, but, as in the former case, the sugar is quickly secreted. Na_2PO_4 produces such rapid secretion that the sugar content of the corpuscles falls temporarily below the normal. Na_2SO_4 causes marked alk. but leaves the secretion of sugar practically unchanged. B. C. A.

Sugar metabolism and water regulation. II. Treatment of cardiac diabetes with insulin and dextrose. H. TATARKA. *Klin Wochschr.* 8, 1763-4(1929), cf C. A. 24,4808.—The action of injected insulin is retarded but more marked; dextrose causes a less rapid and less marked rise in blood sugar. Non-cardiac diabetics with insulin edema behave similarly. B. C. A.

Action of substances of the pilocarpine group (pilocarpine, arecoline, physostigmine) on the gas content of the blood. A. M. LAPOBRAZHIANSKI. *Russkii J. Physiol* 13, 19-32(1930)—Expts. on dogs show that all substances of the pilocarpine group diminish greatly the O content of venous blood and leave unchanged or increase slightly that of arterial blood. Decrease of the latter is, however, caused by large doses of physostigmine. If the decrease of the amt. of O in venous, with normal content in arterial blood is the result of increased O absorption by the tissues, and diminution of the O content of arterial blood the result of insufficient arterialization of the blood in the lungs, the increased consumption of O must be a characteristic common to all substances of the group. The arterialization of the blood in the lungs thus compensates for the O consumed, except that, in most of the cases in which physostigmine and all in which toxic doses of the other substances were administered, the O content of the arterial blood falls. B. C. A.

The comparative changes in gastric acidity and urinary reaction after the injection of histamine. MILTON J. MATZNER AND IRVING GRAY. *Arch Internal Med* 47, 202-5(1931)—In tests combining gastric analysis with histamine as a stimulant and estn. of urinary acidity, not over 50% of the cases showing gastric free HCl gave a definite alk. tide. The use of histamine is discouraged. J. B. BROWN

Liver extract, liver ash and iron in the treatment of anemia. CHESTER S. KEEFER, K. K. HUANG AND C. S. YANG. *J. Clin. Investigation* 9, 533-54 (1930), cl. C. A. 24, 3038.—With liver ext. increased hemoglobin regeneration was shown in the nutritional anemia of childhood and in the anemias of dysentery, hookworm and pregnancy. In some cases the effect of the ext. was increased by Fe. Liver ash, equiv. to 300 g. of liver per day, was of little effect. Fe in the form of FeCO_3 was quite effective, usually in proportion to the size of the dose. Often the effect of Fe was greater than that of liver ext. or ash. J. B. BROWN

Cardiac antagonism of pilocarpine and tropine. RAYMOND HAMET. *Compt. rend.* 192, 111-3 (1931).—Ordinarily the injection of tropine in a dog subjected to the action of a large dose of pilocarpine suppresses the cardiac inhibition. Since the same phenomenon is observed if sparteine replaces the tropine, it is not due to the atropine action of tropine. RACHEL BROWN

Influence of hydrogen-ion concentration on the fixation, by adsorption, of cocaine hydrochloride on the nerve fibers. JEAN RÉGNIER AND GUILLAUME VALETTE. *Compt. rend.* 192, 114-6 (1931), cl. C. A. 25, 143.—The adsorption of cocaine-HCl on animal charcoal increases from pH 3 to pH 7.5 by 1.5 times, the increase being very marked above pH 6. With nerve fibers the increase from pH 3 to pH 7.6 is 5 times. Therefore an alk. medium favors the fixation of the anesthetic on the nerve itself. R. B.

Regulation of respiration. NATHAN B. EDY. *Quart. J. Exptl. Physiol.* 20, 313-31 (1930), cl. C. A. 23, 4251.—The effect upon salivary secretion of the intravenous administration of NaHCO_3 , NaOH , NaCl and Na_2SO_4 . In the dog, salivary secretion elicited by continuous administration of pilocarpine is decreased by intravenous injection of NaHCO_3 , Na_2CO_3 or NaOH . The effect produced by NaHCO_3 is proportional to the size of the dose, whether in isotonic or hypertonic soln. NaOH has a greater action than Na_2CO_3 . NaHCO_3 prevents almost completely the augmenting effect upon salivary secretion of a 10 min. period of rebreathing. These alk. agents also increase submaxillary blood vol. flow. Their effect may be due to changes in acid-base equil. in which the H-ion concn. of the interior of the cell may be a major factor. In isotonic soln. NaCl or Na_2SO_4 increases submaxillary secretion and blood-vol. flow. In hypertonic soln. they decrease the former but increase more markedly the latter. The effect of these salts seems to be due to the increase in blood-vol. flow and to the withdrawal of water from the tissues. The effect upon salivary secretion of the intravenous administration of lactic acid, sodium lactate and HCl. The secretion of the submaxillary gland is increased by the intravenous administration of lactic acid and HCl. The increase due to HCl is frequently followed by a decrease below normal. The effect of HCl is the same whether the sympathetic supply to the gland is intact or not. Lactic acid and HCl increase the submaxillary blood-vol. flow. The increase in secretion produced by lactic acid is greater if the vago-sympathetic has been cut. Na lactate decreases salivary secretion if the vago-sympathetic is intact, but increases it if the nerve is cut. It increases submaxillary blood vol. flow. The lactate ion appears to be inhibitory to the central and augmentory to the peripheral mechanism of secretion. Changes in the acid-base equil. of the tissues and blood are an important factor in the effect of lactic acid, HCl and Na lactate upon secretion. Coincident changes in blood-vol. flow probably affect the results, and some additional effect of the lactate ions, such as their use as food material supplying energy for secretion, may be involved. Effect upon salivary secretion of the intravenous administration of NH_4Cl and $(\text{NH}_4)_2\text{CO}_3$. The administration of NH_4Cl and $(\text{NH}_4)_2\text{CO}_3$ produces an initial increase and subsequent decrease in the rate of salivary secretion and in the submaxillary blood vol. flow. One cc. per kg. of $M/6$ NH_4Cl or $(\text{NH}_4)_2\text{CO}_3$ produces transient fall in blood pressure and brief increase in pulmonary ventilation. Five cc. per kg. of $(\text{NH}_4)_2\text{CO}_3$ lowers the blood pressure and depresses pulmonary ventilation. The effects of NH_4Cl and $(\text{NH}_4)_2\text{CO}_3$ on the salivary secretion are probably due to changes in the acid-base equil. and to changes in blood-vol. flow. RACHEL BROWN

The therapy and the prevention of carbon disulfide poisoning. G. B. AUDO-GIANOTTI. *Industria chimica* 5, 1375-80 (1930).—A summary of the symptoms, therapy and prevention of industrial poisoning by CS_2 . A. W. CONTIERI

Experimental studies in the neutralizing action of sodium thiosulfate in atoxyl intoxication. N. KURODA. *Acta dermat.* 13, 289-99. See C. A. 24, 1431. E. J. C.

Sensitization of the skin of guinea pigs to phenylhydrazine. W. JADASSOHN. *Klin. Wochschr.* 9, 551 (1930). E. J. C.

The effect of insulin upon the rate of dialysis of diabetic blood sugar. ISREAL S. KLEINER, HAROLD BIRNKRAUT AND THEODORE ROTHMAN. *Endocrinology* 14, 226-8 (1930).—The admin. of dry potent insulin to humanized blood of a diabetic dog did not

alter the rate of dialysis of sugar from it. This indicates that the pancreatic hormone does not exert its effect by producing a more readily diffusible form of glucose.

H. J. DEUEL, JR.

Antagonism between sulfur and adrenaline. DOMENICO CAMPANACCI *Liverr Alim. Worksche.* 44, 79-81(1931)—The antigonistic effect on blood pressure is produced when 30 mg of colloidal S and 1 mg of adrenaline are injected. D. B. DILL.

Assimilation of intravenously injected hexoses; phosphorus and water metabolism. E. OWSIANY AND M. WIERZUCHOWSKI *Compt. rend. soc. biol.* 103, 415-8(1930). *Physiol. Abstracts* 15, 240-1—The sugar assimilated is estd. as the difference between the amt. injected and the amt. eliminated by the kidney. Of 4 dogs, 2 assimilated glucose best and 2 fructose, 90% of the total in each case. Insulin causes an insignificant increase in glucose assimilation and has no effect on that of fructose. The amt. of inorg. P in blood and urine is lowered during the active assimilation period. The curve of P reduction is not affected by the rate of assimilation. The P curve is raised again after cessation of injection, and the rise is not prevented by insulin. G. G.

Liver and muscle glycogen and blood sugar in starved and phlorhizinized dogs. F. RATHERY, R. KOVRLSKY AND MILLER V. LAURENT *Compt. rend. soc. biol.* 103, 472-4(1930). *Physiol. Abstracts* 15, 240—Starvation alone, even when prolonged, causes only a slight diminution of hepatic and muscular glycogen. Injection of phlorhizin after a period of starvation causes a diminution of hepatic and muscular glycogen. In spite of the almost complete absence of hepatic and muscular glycogen the glucose level may remain normal, although accompanied by a marked glucosuria. The glycogen of the liver or muscles does not seem to be essential for the maintenance of this normal level. G. G.

Influence of insulin on early variations of hepatic and muscular glycogen in the normal and phlorhizinized dog. F. RATHERY, R. KOVRLSKY AND MILLER V. LAURENT *Compt. rend. soc. biol.* 103, 474-5(1930). *Physiol. Abstracts* 15, 240—Phlorhizin in the fasting dog causes marked diminution of hepatic and muscular glycogen. Inject on of insulin now causes rapid restoration of the glycogen of the liver, while that of muscle is still further diminished. The loss of hepatic glycogen is great, and not proportional to the muscular loss. G. G.

Chemical modifications of blood after injection of sodium nitrate. Y. POTRAK. *Compt. rend. soc. biol.* 103, 536-8(1930). *Physiol. Abstracts* 15, 257—Iodine injected in small doses, which are well tolerated over long periods, causes a disturbance of glucose and cholesterol metabolism manifested by instability of these substances in the blood. Later anemia sets in and is accompanied by a considerable diminution of plasma globulin and an increase of the albumin/globulin quotient. G. G.

Action of choline and its derivatives on respiration. M. VILLARET *Compt. rend. soc. biol.* 103, 675-7(1930). *Physiol. Abstracts* 15, 226—Choline, acetylcholine and formylcholine, which have similar cardiovascular actions, also produce the same effect on the respiratory rhythm. Bromocholine causes more sustained apnea than the others, because it persists longer in the circulation. The effect is due to the choline group. G. G.

Choline apnea and artificial respiration. M. VILLARET *Compt. rend. soc. biol.* 103, 677-9(1930). *Physiol. Abstracts* 15, 226—Choline and its derivs. cause apnea in large doses, polypnea in small doses. The effect is independent of vagus influence. The duration of the apnea is proportional to the dose. Artificial respiration saves the lives of animals treated by otherwise fatal doses. It must be continued much longer with bromocholine than with other derivs. The fatal action of the poisons is effected, not directly on the heart, but on the respiratory center. G. G.

Adrenaline, atropine and lobeline in choline apnea. M. VILLARET. *Compt. rend. soc. biol.* 103, 771-2(1930). *Physiol. Abstracts* 15, 226-7—When apnea is set up by the action of choline, it is not modified by doses of adrenaline, atropine or lobeline, which are capable of counteracting the cardiac influence of choline and its derivs. G. G.

Derivatives of tholine mechanism of polypnea caused by intravenous injection of choline and its derivatives. M. VILLARET *Compt. rend. soc. biol.* 103, 772-3(1930). *Physiol. Abstracts* 15, 227—The polypnea elicited by the injection of choline and its derivs. into the circulation is enhanced by a discharge of adrenaline. The results correspond with the observation that in reflex or central polypnea injection of adrenaline causes, not apnea, but hyperpolypnea. G. G.

Fatty bodies in cells of the liver, lung, kidney, suprarenal and testicle in the normal dog and in the dog poisoned by polyenedamine. J. SŁOWIŃSKI *Compt. rend. soc. biol.* 103, 811-5(1930). *Physiol. Abstracts* 15, 244—In the normal dog there is a

subpleural chain of cellular elements contg. large amts. of cholesterol esters. This reserve seems to be of great importance in defence against intoxication. Intoxication with tolylenediamine leads to infiltration of the liver cells with fat. After poisoning, the disappearance of the subpleural cells and the diminution of fatty bodies in the suprarenal cortex and in the interstitial cells of the testis indicate a defensive process and a neutralization of the toxic substances. G. G.

Histochemical and histological studies of electric ionotherapy. F. RASZEJA. *Compt rend soc biol* 103, 799-803(1930). *Physiol Abstracts* 15, 213-4.—Normally in the skin of the rabbit Ca is distributed in the subcutaneous tissue and K in the epidermis. When CaCl_2 is ionized on a pod with the passage of an elec. current through it to the skin, the K is pushed into the deeper regions and the Ca ions occupy the surface. Rabbits treated in this way are much less susceptible than controls to the decalcifying effects of Ca free diet. The cond. of the serum is markedly increased after introduction of Li in the above manner, and this when the urine becomes free from Li. The elimination of this element lasts longer than is revealed by the urine. G. G.

The role of the reticulo-endothelial system in the metabolism of fat in the normal dog and in the dog poisoned by tolylenediamine. J. SŁOWIŃSKI. *Compt rend soc biol* 103, 816-8(1930). *Physiol Abstracts* 15, 244.—The reticulo-endothelial system in the tonsils, liver and bone marrow plays a most important part in the metabolism of fat. It is still more important in pathol. conditions like tolylenediamine intoxication. The fatty inclusions in the cells of the system are formed of lipoids and cholesterol esters. The elements contg. these particles in the tonsil are situated deeply in the lymphatic nodules. They are rare under the epithelium. G. G.

Pathology of aniline poisoning. G. ARELLO. *Med. Tratat* 1, 294-6(1930).—*Bull. Hyg* 5, 929.—A. describes the symptoms of acute and chronic poisoning by aniline dyes. Venesection and inhalation of ammonia relieved these symptoms. In most cases of dye poisoning the dye can be detected in the urine, particularly if aniline blue or black is the cause. GEORGE R. GREENBANK

Injury to the digestive tract by copying pencils. I. ILKOFF. *Deut. med. Wochenschr* 56, 1132-3(1930). *Bull. Hyg* 5, 929.—The injury is due to tissue necrosis by gradual soln. of the aniline violet. In a test case a man ate meat contg. a 3-cm. piece of the pencil, ulcers formed in the pylorus. Food contg. crushed pieces of the pencil was given to dogs. All the dogs showed gastric or duodenal inflammation and ulcers formed in a few days. This is due to the fact that the aniline deriv. has an alk. reaction which dissolves the albumin and thus causes destructive penetration whereas acid dyes cause coagulation and hinder penetration. GEORGE R. GREENBANK

Insulin and glycogen. I. Study in the normal dog. F. RATHERY AND R. KOWALSKY. *Ann. physiol. physiochim. biol* 6, 32-72(1930).—In the dog the hepatic vein contains the most glucose, the portal vein least and peripheral vessels have intermediate quantities. Following the administration of insulin, the relative order does not change, but the difference between the hepatic and peripheral veins increases. The liver therefore continues to discharge glucose into the circulation; nor can one postulate an increased avidity of the tissues for glucose, as the difference between the arterial and venous blood sugar does not increase. Immediately following the injection of insulin, the blood sugar shows a transitory increase, particularly in the hepatic vein, the liver glycogen decreases, while the muscle glycogen is not appreciably changed. Conclusions are given in the following paper. II. Study in the dog following pancreatectomy, starvation or phlorhizin. General conclusions concerning the role of glycogen. *Ibid* 73-123.—Following pancreatectomy, all the sugar values rise, but their order changes. The sugar of the hepatic vein falls below that of the peripheral veins, however, it remains higher than the portal blood sugar. Liver glycogen falls precipitously, while the muscle glycogen is only slightly affected. Qualitatively, the effects of insulin are exactly the same as in the normal animal, except for the fact that it induces an increase in liver glycogen. Moreover, insulin has the same effect upon starved dogs as upon normal dogs. Phlorhizin administered to fasting dogs causes an almost complete disappearance of liver glycogen, but the blood sugar remains normal. Conclusion. Liver glycogen plays no role in the maintenance of blood sugar; there is some intermediate product of metabolism in the liver. Insulin does not mobilize glycogen or activate the muscle metabolism of sugar. The effect of insulin upon glycogen, whether hepatic or muscular, is of only secondary importance in insulin hypoglycemia. Glycogen is not indispensable either for the proper functioning of the liver and muscles or for sugar metabolism. H. EAGLE

Behavior of soaps in the animal organism. IRVINE H. PAGE AND E. V. ALLEN. *Arch. Exp. Path. Pharmacol* 152, 1-27(1930).—Most soaps, particularly those of

hydroxy and unsatd acids, are highly toxic. Injected intraperitoneally into rats and mice, they cause marked histological changes, chiefly perivascular infiltration and focal degeneration in the liver and proliferative changes and adhesions in the peritoneal cavity. The total fat content of the liver decreases, but the I no increases. On intra venous injection of soap, the blood pressure falls and respiration increases. Diethanol amine ricinoleate increases the permeability of the blood spinal fluid barrier for circulating dyes. H EAGLE

The action of a sympathomimetic alkaloid in *Sida cordifolia* (Breia). R. N. CHOPRA AND PREMANKUR DA *Indian J. Med. Research* 18, 467-75(1930)—The active principle of *Sida cordifolia* (Breia) is a sympathomimetic alkaloid resembling ephedrine in its pharmacol properties. H EAGLE

The conditions of "basic narcosis" in combined narcosis experiments. L. LEVDE *Klin Wochschr* 9, 1609-15(1930) H EAGLE

Secondary ultra-violet radiation and its biological effect. MONA SPIEGEL-ADOLF *Klin Wochschr* 9, 1615-8(1930)—Only a small portion of the radiation-energy emitted by Ra or x rays is absorbed by proteins in soln., while ultra-violet light is comparatively much better utilized. The fluorescent light emitted by salt crystals following exposure to radium or x rays is much more effective in causing the coagulation of protein, the hemolysis of red cells and the death of parameria than the original radiation energy used to induce this fluorescence. H EAGLE

The effect of thyroxine. O. EICHLEA AND R. SANDESS. *Klin. Wochschr.* 9, 1618(1930)—The intravenous injection of thyroxine (1.8 mg per kg) into rabbits causes a significant increase of the blood lactic acid and O_2 consumption. H E

The antagonism between adrenaline and insulin. Experiments in adrenalectomized dogs. R. CARIO *Klin Wochschr* 9, 1623-4(1930)—The injection of conod. glucose solns into the pancreaticoduodenal artery of dogs does not cause an elevation of blood sugar even after the adrenals have been removed. Injected into any other artery, it causes a transient hyperglucemia. This is further evidence that the blood sugar as such is an adequate hormone for the regulation of insulin secretion. Adrenaline and insulin may be antagonists in the regulation of blood sugar. H EAGLE

Animal and clinical studies on the question of synthalin-B. HERBERT HIRSCH KAUFMANN *Klin Wochschr* 9, 1631-2(1930) H EAGLE

Oral desensitization of dermatoses due to alimentary idiosyncrasies by the use of species-specific peptones. ERICH URBACH. *Klin Wochschr* 9, 2046-9(1930) H EAGLE

The opposing effects of liver and spleen. A contribution to the spleen treatment of erythremia. F. HÖGLER. *Klin Wochschr.* 9, 2052-8(1930)—Ra irradiation of the long bones has a beneficial effect in erythremia, in which it is often followed by an excretion of aldehyde in the urine, but no beneficial effect could be discerned upon pernicious anemia. In 2 cases of erythremia a striking improvement in the blood picture was effected by the daily ingestion of large quantities of spleen over a period of months. In one case of pernicious anemia, splenectomy followed by liver therapy caused a typical erythremia, which was controlled by the subsequent administration of spleen. The liver and spleen therefore have antagonistic effects upon the erythrocytes. H EAGLE

The separation of mixtures of dyes by normal and poisoned kidneys. RUDOLF HÖBER. *Klin Wochschr* 9, 2065(1930), cf C A 24, 4543—Most dyes are excreted slowly through the glomeruli (e g., cyanol) but some, like phenol red, are rapidly excreted by the tubules in a conod. form. If a mixt. is injected into frogs or rabbits, the tubular dye predominates at first, after it has been almost completely excreted, the urine contains only the glomerular dye. If the tubules are injured, as by uranium nitrate, the quick tubular excretion of a conod. dye is not observed, and one obtains only the glomerular excretion of the two in approx. the same concn. This fact suggests a method for testing tubular function in man. H EAGLE

A highly active liver extract which can be injected. M. GANSSLEY. *Klin Wochschr* 9, 2099-103(1930)—Excellent results were obtained in the treatment of pernicious anemia by the intragluteal injection of protein free liver ext. Improvement is usually obtained in 6-8 weeks. H EAGLE

Clinical observations on the treatment of anemias with stomach preparations. FERDINAND BERTRAM *Klin Wochschr* 9, 2103-6(1930)—Using the proprietary prepn. Stomopson, B found improvement in 10 of 11 cases of Biermer's (pernicious) anemia. In the unsuccessful case liver therapy was effective. H EAGLE

Experimental studies on the effect of the inspiration of small quantities of benzine and benzene upon the respiratory organs and the entire body. M. SCHMIDTMANN.

Klin. Wochschr. 9, 2100-8(1930).—The inhalation by rabbits, guinea pigs, mice a rats, of small quantities of benzene or benzene-gasoline vapor over long periods of time eventually causes emphysema, chronic bronchitis, and atelectatic pneumonia. Initial leucocytosis is followed in the later stages by a marked leucopenia, and a fall in both the red cell count and hemoglobin. In the spleen, an initial myeloid infiltration of the pulp is followed by a gradual decrease in cells, pronounced phagocytosis of red cells, and, in some cases, hyalinization. These results are exactly the same as the changes induced by the subcutaneous injection of benzene, differing only in the time required for their evolution (months). Heart and liver both show focal necrosis in the center of small areas of infiltration. The cumulative toxic effect of these repeated small inhalations are highly significant in the light of the continuous exposure of urban population to the exhaust of automobile engines. H. EAGLE

The effect of beef adrenal extract upon the uterus. ERICH ENGELHART. *Klin. Wochschr.* 9, 2114-5(1930).—The subcutaneous injection of beef adrenal extract in virgin female rabbits causes proliferative changes in both the mucosa and musculature. H. EAGLE

The decrease of oxidation processes during ether anesthesia. H. FUSS AND DERRA. *Klin. Wochschr.* 9, 2115-6(1930).—In dogs ether drop anesthesia causes definite arterial O_2 -unsaturation, which is less pronounced with ether air anesthesia, and is observed if an ether- O_2 mixture is used. The increase in blood lactic acid during anesthesia bears some relation to the degree of arterial O_2 -unsaturation and is possibly due to an inhibition of resynthesis. H. EAGLE

Copper treatment of anemia in nurslings. E. SCHIFF, H. ELIASBERG AND JOFFE. *Klin. Wochschr.* 9, 2144-5(1930).—The authors believe that the daily administration of 20 drops of 1% $CuSO_4 \cdot 5H_2O$ is of benefit in anemia. H. EAGLE

The treatment of pernicious anemia with preparations of gastric mucosa. HENNING AND G. STIEGER. *Klin. Wochschr.* 9, 2145-7(1930).—A dried prep. of gastric mucosa was found to be highly effective. Acidic extracts were not effective. Both the antrum and fundus contain the active principle, which is not pepsin. H. EAGLE

A case of hydrazoic acid poisoning. Zs. KÖCHER. *Klin. Wochschr.* 9, 2166(1930). H. EAGLE

The effect of intracutaneous injection upon the basal metabolism. WILFRIED LOWENSTEIN. *Klin. Wochschr.* 9, 2255-6(1930).—The intracutaneous injection of solutions of protein, or even H_2O , causes a marked fall in the basal metabolic rate, amounting to as much as 50%, and beginning within 10-15 min. L. suggests that it is due either to a reflex in the vegetative nervous system, or to the release of histamine from cells at the site of injection. H. EAGLE

The significance of ammonium chloride medication in the treatment of diuretics. PAUL SAXL AND OTTO ERLSACHER. *Klin. Wochschr.* 9, 2302-3(1930).—NH₄Cl causes a lowered alkali reserve and decreased circulatory blood volume, while Salicylic acid causes an increased blood volume. Despite their antagonistic action upon the volume of circulating blood, a combination therapy with these 2 drugs causes a much more rapid removal of transudates due to cardiac failure, hepatic cirrhosis, nephritis, etc., and of exudates, as in articular arthritis. H. EAGLE

The dependence of gas exchange and the action of iodine upon the hydrogen-ion concentration. I. WISLICKI. *Klin. Wochschr.* 9, 2354-5(1930).—Increased [NaHCO₃] or decreased CO₂ tension causes an increased O_2 consumption by rat muscle *in vitro* and an increased respiratory quotient. An acid reaction has the opposite effect. NH_4I in small concn. neutralizes either change, in higher concn. it causes an increased formation of lactic acid without any further effect upon O_2 consumption. It is interesting to note that it has no constant effect at serum reaction (pH 7.46). Thyroxine has no effect, probably due to its insolubility at this H -ion concn. H. EAGLE

Excretion and storage of neodorm. ANNELEISE ROTHMAN. *Klin. Wochschr.* 9, 2398-9(1930).—Weeks after the ingestion of neodorm (bromoisopropylbutyramide), Br₂ can be demonstrated in the tissues, the brain in particular, in a water sol. form. The Br₂ is gradually excreted, chiefly in the urine and bile. H. EAGLE

Pharmacological study of the effect of caffeine-adenine mixtures. DAVID MACHT AND HERMANN SCHROEDER. *Klin. Wochschr.* 9, 2429-30(1930). H. EAGLE

Treatment of insulin lipodystrophy. REINHOLD BOLLER. *Klin. Wochschr.* 9, 2433-5(1930).—The addition of cocaine or novocaine in insulin has a preventive therapeutic effect in diabetics who tend to develop lipodystrophy as a complication of the insulin treatment. H. EAGLE

The effect of digitalis upon the final wave of the electrocardiogram. ERICH BLUMENFELDT AND SPENCER G. STRAUSS. *Klin. Wochschr.* 9, 2439(1930).—?

reversal of the T-wave in the electrocardiogram is only an exceptional occurrence in patients receiving digitalis and cannot be taken as a criterion of digitalis effect.

H. EAGLE

The effect of vasopressin and oxytocin upon the blood sugar in human beings. A. W. L. ELMEZ AND M. SCHIFF. *Klin. Wochschr.* 9, 2439-40 (1930).—Vasopressin, injected subcutaneously, usually causes an elevation in blood sugar. The effect of pituitrin is much less frequent and less marked. Oxytocin is ineffective. H. EAGLE

Local anesthetics. Derivatives of amino alcohols having the function of primary alcohols. E. FOURNEAU, G. BENOIT AND ROGER FIRMENICH. *Bull. soc. chim.* 47, 58-55 (1930).—The prepn. and properties of a series of di primary glycols are described, also the aldehydes from which they are derived, several of which are new. Similar tables are given of a series of acetobromohydrins, and a series of amino alcs. with their benzoyl derivs. Some of these derivs. which formed colorless crystals of definite m. p. were tested as local anesthetics. They are very sol. in H_2O and alc., sol. in acetone, insol. in Et_2O and taste bitter. They have a strong anesthetic reaction which takes effect slowly.

ANN NICHOLSON HIND

Carbon monoxide poisoning and a new treatment by irradiation. F. KOLA. *Rechtskr. Lekárske Listy* 9, 16-33 (1929).—Ultra-violet light causes disson. of CO hemoglobin *in vivo* and *in vitro*. Animals and humans poisoned with CO showed marked improvement when irradiated with ultra-violet light at a distance of 60 cm. for 40 min.

WILLIAM J. HERRA

Influence of arsenicals and crystalline glutathione on the oxygen consumption of tissues. CARL VONGLIN, SANFORD M. ROSENTHAL AND J. M. JOHNSON. *U. S. Pub. Health Repts.* 46, 339-54 (1931).—The addn. of cryst. SH-glutathione to kidney, liver, testis, the Jensen rat sarcoma, and baker's yeast does not increase the rate of O_2 consumption beyond the extra amt. of O_2 required to oxidize the S of the added glutathione. Oxidized glutathione has no accelerating influence on the O_2 consumption of kidney and testis. Arsenous oxides ($RAsO_3$) in relatively low concn. cause a pronounced reduction in the rate of O_2 consumption. The pentavalent arsenicals ($RAsO_5H_2$), including trypanamide, in the same concn. are devoid of any influence on the O_2 consumption. Of the arsenobenzene derivs. ($RAs=AsR$) sulfarsphenamine is ineffective, whereas neararsphenamine due to its rapid oxidation reduces the O_2 consumption, but less markedly than arsenoxide. These results are in harmony with observations concerning the pharmacol. and chemotherapeutic properties of these compds. which distinguish the 3 groups, $RAsO_3$, $RAsO_5H_2$, and $RAs=AsR$. SH-glutathione when added to tissues in the ratio of 10 moles to 1 mole of arsenoxide prevents the reduction in O_2 consumption caused by arsenoxide alone. S-S-glutathione is ineffective, showing that the action of SH glutathione is due to its SH group. $FeSO_4(NH_4)_2SO_4 \cdot 6H_2O$ is ineffective in overcoming the reduction in O_2 consumption produced by arsenoxide. These observations add further evidence in favor of the theory that the pharmacol. action of these arsenicals is essentially due to a chem. reaction with SH glutathione and possibly other SH compds. of protoplasm. From the physiol. viewpoint the results appear to indicate that glutathione in some as yet unexplained manner is concerned in the O_2 consumption of tissues *in vivo*. J. A. KENNEDY

Observations on the venom of a large Australian snake, *Pseudechis australis*. C. H. KELLAWAY AND DONALD F. THOMSON. *Australian J. Exptl. Biol. Med. Sci.* 7, 134-50 (1930), cf. *C. A.* 24, 650.—Weekly "milks" through rubber yielded 0.1 to 0.31 g. dry venom. It exerted a marked effect upon heart muscle, blood pressure and respiration, and is strongly hemolytic and anticoagulant. The lethal dose for lab. animals varied from 0.6 to 5 mg. per kg.

C. G. KING

The action on mammalian circulation of the skin secretion of *Hyla aurea*. W. A. OSBORNE. *Australian J. Exptl. Biol. Med. Sci.* 7, 226-8 (1930).—The secretion exerts an effect like that of digitalis on the heart and like that of ephedrine on the arterioles.

C. G. KING

Standardization of the trypanocidal activity of some aryl derivatives of arsenic and L. LATOY AND (MISS) ENGLE. *Bull. soc. chim. ind.* 12, 886-93 (1930).—A new standard index is proposed, based upon "trivalent curative dose" (I) and "tolerance dose" (II), in which I is the av. quantity required to destroy a virulent blood stream infection of *Tr. brucei*, *Tr. evansi* and *Tr. equiperdum* in mice (held 30 days) and II is the intravenous dose tolerated by 50 to 60% of the animals. Rabbits may also be used. The index so obtained for atoxyl is taken as one and other compds. are given proportionate values, trypanamide = 2.03 and Fournau 270 = 5.95.

C. G. KING

Iodine reaction in xerophthalmic goiter. WILLARD G. THOMPSON AND PHILIP K. THOMPSON. *Endocrinology* 14, 373-9 (1930); cf. *C. A.* 24, 3559.—In Boston the small-

est quantity of I that produces any reduction in basal metabolism in exophthalmic goiter is greater than 0.75 mg in $\frac{1}{2}$ of the cases. The min. amt. to produce max. reduction is 6 mg. Between these limits, the relation between the percentage approach of the basal metabolism toward the normal level and the size of the dose is roughly linear. The effective dose has not been detd in Chicago but 6 mg causes a marked reduction in basal metabolism.

MARY E LEAR

The role of iodine in the treatment of exophthalmic goiter. M. L. GREENSTEIN. *U. S. Vet. Bur. Med. Bull.* 7, 167-9(1931)—G describes the beneficial effects of Lugol's soln. in a specific case.

B. S. LEVINE

Experimental rabies in white mice and attempted chemotherapy. A. HOYT AND C. W. JUNGEBLUT. *J. Infectious Diseases* 47, 418-24(1930)—Prophylactic administration of various drugs (particularly arsenicals and quinine derivs) failed to protect white mice infected intracerebrally by single or multiple minimal lethal doses. The period of incubation of the disease consistently showed a very slight prolongation following injection of Ag arspenamine.

JULIAN H. LEWIS

The effect of barium sulfate upon the incidence of human intestinal protozoa. JUSTIN ANDREWS AND MOSES PAULSON. *J. Lab. Clin. Med.* 16, 39-42(1930)—The no. of intestinal protozoa per unit vol. of stool appears to be greatly reduced following the ingestion of BaSO_4 . The intestinal amebas are affected more noticeably than the flagellates. The effect is transitory and is not prolonged by extension of the period of administration of BaSO_4 .

E. R. MAIN

Pharmacology of brominated oils. CLAYTON S. SMITH AND HELEN L. WIKOFF. *J. Lab. Clin. Med.* 16, 43-6(1930)—The brominated esters of cottonseed oil appear to be well adapted for use in the roentgenologic demonstration of lesions. They are eliminated from the site of injection in rabbits within a week and produce no deleterious effects in repeated doses.

E. R. MAIN

Antimony in medicine. C. N. MYERS AND BINFORD THRONE. *J. Lab. Clin. Med.* 16, 239-46(1930)—A compd. (M 303) is described which contains both Sb and As in the trivalent condition. It appears to be a valuable therapeutic agent in syphilis, leprosy and diseases in which protozoal organisms are involved. An historical discussion of the use of Sb in medicine is included.

E. R. MAIN

The elimination of water and sodium chloride through the kidneys after intravenous injection of isotonic sodium chloride solution. IGNAZIO SIRONI. *Biochim. terap. sper.* 17, 469-77(1930)—A soln. of 0.974% NaCl was slowly injected into rabbits. In proportion to the vol. injected a higher percentage of Cl and H_2O is eliminated, the ratio of $\text{H}_2\text{O}:\text{Cl}$ is always the same.

A. E. MEYER

A new digitalis, "Digitalis lanata" Ehrh. EM. PERROT, P. BOURCET AND RAYMOND-HAMET. *Bull. sci. pharmacol.* 38, 7-16(1931)—The plant, whose botanical description is given, contains a cryst. glucoside, dilamin. of digitalis action, but of higher potency than the cryst. digitalin of *Digitalis purpurea*. The toxicity is the same in oral administration as in intravenous injection. In contrast to *D. purpurea*, the *D. lanata* keeps up an almost const. pharmacol. activity when cultivated. The drug is not only more uniform, but also of higher content of the glucoside.

A. E. MEYER

Quinidine in paroxysmal tachycardia. C. BONORINO UDAONDO. *Semana méd.* (Buenos Aires) 1930, 11, 1402-7—Quinidine sulfate (0.5 g) dissolved in physiol. saline was given intravenously. Two cases were successfully treated.

A. E. MEYER

Sinusitis treatment with Lipiodol. RICARDO BRACHT. *Semana méd.* (Buenos Aires) 1931, I, 102-4.—Iodized oil, Lipiodol, was injected into the sinus cavity for radiography. It is an excellent remedy.

A. E. MEYER

Local inflammations treated with concentrated solution of magnesium sulfate. S. JOUAN. *Semana méd.* (Buenos Aires) 1931, I, 259-62.—A satd soln. of MgSO_4 was applied in a permanent bandage on the inflamed area. Forty cases were successfully treated.

A. E. MEYER

Influence of insulin on the glycogen content of the perfused rabbit liver. NIELS A. NIELSEN. *Biochem. Z.* 230, 259-68(1931)—On perfusing the rabbit liver with defibrinated blood without insulin there is a gradual decrease in the sugar content of the blood and a corresponding increase in the liver glycogen. The deposition of the glycogen in the liver is not uniform. Addn. of $\frac{1}{2}$ to 20 units of insulin to the blood causes a rise in the blood-sugar concn. and a loss of glycogen in the liver. No effect was observed with amts. smaller than $\frac{1}{2}$ insulin units.

S. MORGULIS

Influence of fermentation poisons upon experimental tumors. I. L. KARCAZAG. *Biochem. Z.* 230, 411-9(1931)—Enzyme and protoplasm poisons have been differentiated upon the basis of their effect in inhibiting the fermentation of (1) both yeast and zymine (protoplasm and enzyme poison), (2) zymine alone (enzyme poison), or (3)

yeast alone (protoplasm poison). By this method it was possible to discover substances which injured the glycolytic enzyme without producing any poisoning effect upon the protoplasm. The N of the biologically active mol plays a paramount role, especially the NO_2 and the CN groups, and not in the nucleus but in the side chains of the compd. By the same method substances were discovered which are protoplasmic poisons. It was then attempted to select substances such as isoquinoline red, Magdal red, ethyl red, Martius yellow, Mikado yellow, bromothymol blue, propionaldehyde, cyanohydrin, nitrobenzaldehyde, o-nitrobenzonitrile, which because of their relative nontoxicity could be employed in animal expts to treat tumors. No consistent results were obtained, although with isoquinoline red, Magdal red and ethyl red a definite influence was noted. II. L. KARCIAG AND C. SEILER *Ibid* 420-34.—Ethyl red, chrysoidine, Magdal red, isoquinoline red and Martius yellow, which are inhibitors of the yeast fermentation process, were found to exert an inhibiting action on the glycolysis of tumor cells which diminishes in the order given above. These substances also cause a diminution of the respiration in the same order so that Martius yellow either had no influence upon the respiration of the tumor cells or actually stimulated it. These enzyme poisons act upon several enzymes so that both the fermentation and the respiration enzyme are simultaneously inhibited, and in this respect they differ from the KCN which injures the respiratory enzyme much more than the glycolytic enzyme. Furthermore, whereas the action of KCN on glycolysis is reversible, the effect of the above substances is irreversible. *In vivo* the effect of these enzyme poisons can be superimposed upon the action of KCN, but when used subcutaneously in animals the KCN effect predominates. The substances which were found to be toxic for the fermentative activity of both tumor and yeast cells were also toxic for the fermentation by *B. coli*. It is usually easier to influence glycolysis than respiration, and it is therefore concluded that the therapeutic efforts, so far as the metabolism of the tumor is concerned, must be directed toward the respiratory function. It is suggested that results could only be expected by definitely and permanently inhibiting their respiration while leaving the glycolytic ability unchanged or slightly lowered. III. L. KARCIAG *Ibid* 435-8.—The sugar content of Ringer soln mixed with sarcoma powder increases. Fermentation of the glucose proceeds slowly and continuously in such a mixt., though the controls show no fermentation in the same time. Conclusion: Rous sarcoma can ferment glucose with liberation of gas and this process can be studied also *in vitro*.

S. MORGULIS

The action of certain autonomic drugs upon the pigmentary responses of *Fundulus*. DIETRICH C. SMITH *J. Exptl. Zool.* 58, 423-53 (1931).—Certain sympathetic and parasympathetic stimulants (cocaine, ergot, pilocarpine, physostigmine, atropine) were injected into the body cavity of the fish, *Fundulus heteroclitus*, and the effects upon the dermal melanophores were studied. Cocaine in sublethal doses produces a contraction, in lethal doses an expansion of the melanophores. Ergot produces a stellate condition of the melanophores which then respond to temp. changes like denervated melanophores. Pilocarpine and physostigmine retard the normal responses of the melanophores. High concns. produce permanent expansion. Atropine produces a stellate condition but unlike pilocarpine the melanophores still respond to background shade. Lethal doses produce complete expansion but the melanophores respond to temp. changes like innervated melanophores. In appropriate dosage atropine antagonizes the action of pilocarpine but not of cocaine. The melanophores of fishes possess a double rather than a single innervation. C. H. RICHARDSON

Adsorption and narcotic action. H. H. KING, J. LOWE HALL, A. C. ANDREWS AND H. L. COLE *J. Pharmacol.* 40, 275-89 (1930).—The authors detd. the free surface tensions of drugs in aq. solns., the interfacial tensions of aq. solns. against pure liquid paraffin and against a liquid paraffin soln. of lecithin, and called the adsorption at various concns. The drugs used were: trional, butyl chloral hydrate, sulfonal, bromal hydrate, triacetin, diacetin, chloral hydrate, Et carbamate, monoacetin, Me carbamate, Et Me carbamate. A comparison of the compds. in homologous series showed that the free surface tension and the adsorption in each series decreased as the liminal values (as detd. by Meyer) increased. The calcd. adsorption values increased as the liminal values decreased in the series with pure paraffin. For the paraffin-lecithin series the results were not so regular. C. RIEGEL

Chloroform content in various tissues during anesthesia and its relationship to the theories of narcosis. JAMES L. MCCOLLUM *J. Pharmacol.* 40, 305-25 (1930).—Dogs were anesthetized with CHCl_3 and the blood and tissues analyzed for CHCl_3 content. The corpuscles contained more CHCl_3 than the plasma and more than their lipid content would indicate. The CHCl_3 content of the brain and tissues increased grad-

ually, but at the point of highest concn (for the brain) there was less CHCl_3 than the lipid content would indicate was possible. C. RIEGEL

Number of open glomeruli in acute mercuric chloride nephrosis. ROBERT A. MOORE AND LOUIS M. HELLMAN. *J. Exptl Med* 53, 303-6 (1931).—Acute lig nephrosis in the rabbit is not associated with a decrease of glomerular circulation. C. J. W.

Effects of the intravenous injection of colloidal silver upon the hematopoietic system in dogs. SAMUEL S. SHOUSE AND GEORGE M. WHIFFLE. *J. Exptl Med* 53, 413-20 (1931).—Colloidal Ag has no sp. action on the bone marrow in dogs but is a systemic poison which may cause anorexia, weakness, loss of wt., anemia and death. Hemolysis can be demonstrated after large doses of colloidal Ag and the anemia presumably is due in part at least to a destruction of red blood cells in the peripheral circulation. The colloidal Ag, injected intravenously, is deposited as granules almost exclusively in the cells of the reticul-endothelial system after the manner of particulate substances. Repeated injections of non-lethal amts. of this substance are invariably followed by hyperplasia of the bone marrow. Large single doses cause rapid death in 12 hrs. or less, characterized by pulmonary edema and congestion. C. J. WEST

Aplasia of marrow and fatal intoxication in dogs produced by Röntgen radiation of all bones. SAMUEL S. SHOUSE, STAFFORD L. WARREN AND GEORGE M. WHIFFLE. *J. Exptl Med* 53, 421-35 (1931). C. J. WEST

The combined effects of colloidal silver and highly filtered Röntgen radiation upon the hematopoietic system in dogs. SAMUEL S. SHOUSE AND STAFFORD L. WARREN. *J. Exptl Med* 53, 437-45 (1931).—The individual destructive effects of colloidal Ag and heavily filtered radiation are still evident when the 2 are used together; the combined effects are cumulative in that small doses are more destructive than when either is used alone. The leucocytosis resulting from the injection of the colloidal Ag affords no protection against the terminal leucopenia following the radiation. C. J. W.

The effect of amyl nitrite on the size of the heart and the width of the aortic shadow as determined roentgenologically. WM. A. BRAMS AND HERMAN A. STRAUSS. *Am. J. Med. Sci.* 180, 618-20 (1930).—Six normal and 15 arteriosclerotic persons were studied roentgenologically before and after administration of AmNO_2 . The transverse diam. of the heart shadow was reduced in 18 of the subjects, the presence or absence of arteriosclerosis seemed to play a minor role in these changes. The width of the aortic shadow was increased in 13 of the 21 patients after AmNO_2 administration. This result was independent of the condition of the peripheral vessels or the degree of blood pressure. R. C. WILLSON

The action of thymophysin on the human pregnant uterus in situ. M. PRINCE RUCKER. *Am. J. Obstet. & Gynec.* 20, 791-7 (1930).—After an intramuscular injection of 0.5 cc. of thymophysin in a case at term there was apparently no change in the type of uterine contractions or in the tone of the uterus. A dose of 1 cc. in a case in the fifth month of pregnancy showed an increased frequency, a decreased force of contractions and an increase in the tone of the uterus between contractions. The same dose produced similar but more marked changes in cases at term. R. C. WILLSON

A study of the effect of acriflavine given intravenously on experimental uterine infection in the dog. JOSEPH L. MEYER. *Am. J. Obstet. & Gynec.* 20, 760-74 (1930).—An appropriate dose of acriflavine given intravenously has a definite beneficial effect on the course of the experimentally infected uterus of the dog. The earlier the dye is injected the greater is the effect. In some instances sterile cultures of the inoculated or infected uterine horn are obtained after acriflavine therapy. R. C. WILLSON

Cerebral blood flow. I. The effect of intravenous administration of hypertonic and hypotonic solutions on the volume flow of blood through the brain. CONN. PILCHER. *Arch. Neurol. & Psychiatry* 24, 899-906 (1930).—The O content of arterial blood and blood from the occipitovertebral sinuses in the dog was studied before and after the intravenous injection of 50% dextrose or 30% NaCl and of distd. water. The O content varied immediately after the injections but became stable within 30 min. and remained so for several hrs. The O content from the occipitovertebral sinus was invariably diminished markedly after the injection of both hypotonic and hypertonic solns. The arterio-venous difference (utilization of O) increased after injection of both types of solns. II. The effect of intravenous injection of hypertonic and hypotonic solutions on the cardiac output and blood pressure. *Ibid.* 907-12.—The cardiac output was usually diminished after injection of both types of solns. The O consumption and O utilization were usually increased by both types. The mean blood pressure varied during the injection but thereafter was increased slightly in all expts. R. C. WILLSON

The effect of potassium iodide on the pulse rates of normal individuals. DONALD

McEACHERN *Bull Johns Hopkins Hosp* 47, 290-303(1930) —In the absence of thyroid disease or other interfering circumstances, KI produced significant increases in pulse rate in 12 subjects while S showed a slight decrease. The changes came on gradually about the third to seventh day and reached the max at about the tenth day. Normal rate was restored within a week or so after the drug was discontinued.

R. C. WILLSON

The effect of the ingestion of potassium iodide on the electrocardiogram of normal individuals. DONALD McEACHERN AND B. M. BAKER. *Bull Johns Hopkins Hosp* 47, 304-7(1930) —KI in therapeutic doses has no significant effect. R. C. WILLSON

The effect of lead on the isolated uterus. M. D. DEVLET KILDERVA AND B. A. RAIKHEIN. *Kazansky Meditsinsky Journal* 26, 690-4(1930). *J. Am. Med. Assoc.* 95, 1545 —After a graphic record had been made of the normal muscular contractions of a part of the uterine cornu in cats placed in Ringer's or Locke's soln kept at 35° with a const. stream of O flowing through the soln, Pb acetate was added gradually to the soln. The uterine contractions changed with the concn of the Pb acetate. When this reached 1/200 the uterus reacted by a complete tetanus and the muscular contractions could not be made to reappear even after all traces of the Pb acetate had been washed out. The introduction of Pb acetate into the blood stream of cats also produced a hypertonic condition of the uterus.

R. C. WILLSON

Antagonism between adrenaline and ergotamine, and the respiratory quotient. R. CAPO. *Riforma Medica* 46, 1347-53(1930). *J. Am. Med. Assoc.* 95, 1703 —Ergotamine does not modify the greater consumption of O induced by adrenaline but when administered during fasting it causes a notable lowering of the elimination of CO₂ and a reduction of the R.Q. Adrenaline causes a mobilization of hepatic glycogen and thus increases almost constantly the elimination of CO₂ and rises the R.Q. Ergotamine inhibits the hyperglucemia that is induced by adrenaline without modifying the increased intra-org. combustion caused by the latter.

R. C. WILLSON

Iodoxybenzoate as a test reagent for free phenolic hydroxyl groups in organic compounds (LEACH) 7. Trypsin preparations suitable for the prevention of adhesions (WALTON) 17. Derivatives of anesthenn (COVICILIO) 10.

ANSCHÜTZ, W., SPECHT, K. AND TIEMANN, FR.: Die Avertunnarkose in der Chirurgie. Berlin J. Springer 200 pp. Reviewed in *J. Am. Med. Assoc.* 96, 715(1931)

I-ZOOLOGY

K. A. GORTNER

Amylase in the extracts of the salivary glands of the silk-worm larva (*Bombyx*). S. MATSUMURA. *Bull Sericulture and Silk Ind (Japan)* 2, 4-5(1930) —An glycerol exts. of the salivary glands of the silkworm hydrolyze sol. starch to dextrin and sugar and glycogen to dextrose with optimum activity at pH 6.6-7.1. In the absence of NaCl the enzyme is almost inactive. Exts. made from European varieties of silkworm are less active than similar preps. from the Chinese and Japanese varieties.

B. C. A.

Cardiac hormone in the tortoise. A. DE CARVALHO. *Compt. rend. soc. biol.* 102, 535-37(1929) —Elec. irritation, and the action of pilocarpine, atropine, ephedrine and ergotamine on the heart of a tortoise failed to detect a cardiac hormone.

B. C. A.

Action of hypophyseal extract on egg production of batrachians. R. KERN. *Compt. rend. soc. biol.* 103, 744-5(1930); *Physiol. Abstracts* 15, 248 —The alk. aq. ext. of the anterior lobe of the hypophysis of mammals, when injected into *Discoglossus*, rapidly elicits the free laying of eggs in the females captured in Dec. and Jan. Controls which are kept in the same conditions of fasting, etc., do not lay their eggs.

G. G.

Cause of ammoniacal odor of flesh of the shark. O. MARTIN. *Z. Fleisch Milch-Hygiene* 40, 48-9(1929) —The NH₃ arises from urea (up to 2.5% of the flesh), the origin of which is not pathol., it serves to maintain the internal osmotic pressure against that of the water.

B. C. A.

Amino acids of tissues. IV. The diamino acid content of muscle tissue of different classes of animals. JOHN L. ROSEDALE AND JOSEPH P. MORRIS. *Biochem. J.* 24, 1294-6(1930). cf. C. A. 24, 161. —An increase in the amt. of histidine N and a decrease in lysine N are found in the lower organisms (tissue of crocodile, cockroach, mosquito larvae, oysters and sponges). V. Monoamino acids of the horse mackerel (*Caranx rottileri*). *Ibid* 24, 1297-1300 —The monoamino acids are extd. with NaOH. Better results are obtained by first removing some of the amino acids from soln.

BENJAMIN HARROW

The hydrolytic products of the protein from the egg capsule of the gastropod, *Hemifusus tuba* Gmel. JUN-ICHIRO SAGARA. *J. Biochem. (Japan)* 12, 473-4(1930).—The protein of the egg capsule of *Hemifusus* seems to belong to the albumoids, being in some respects similar to keratin and in other respects to elastin. It contains no glycine, tyrosine or histidine, but much leucine and glutamic acid, these 2 amino acids making up 30% of the protein. S. MORIGUTI

A qualitative analysis of the digestive secretions of the larva of the Japanese beetle (*Popillia japonica* Newm) MILLARD C. SWINGLE. *J. Econ. Entom.* 23, 956-8 (1930).—A qual chem examn of the digestive secretions of this larva showed the presence of ions of Al, Fe⁺⁺, Ca, Mg, K, Na, NH₄, CO₃, Cl, NO₃ and PO₄. C. H. R.

Investigations on the rapid coloration of the cilia of ciliates. HELENE ANSELMIER. *Pharm. Acta Helv.* 5, 33-41, 64-71, 88-94(1930).—Detailed directions are given for the differential staining of the cilia of ciliates, thus making it possible rapidly to diagnose ciliates in mixed cultures of infusoria. Photomicrographs of certain protozoa of various origins, and a bibliography of 10 references accompany this highly specialized study. S. WALDBOTT

12—FOODS

F. C. BLANCE AND H. A. LETTER

Amendment of July 8, 1930, to the Federal food and drugs act and requirements thereunder. ANON. Food and Drug Administration, U. S. Dept. Agr., *Service and Regulatory Announcements* F. D. No. 4(1931).—The forms of statements required on canned foods of substandard quality and substandard fill, together with standards for fill of container of all canned foods and of quality and condition for peas, peaches and pears, are given. H. A. LETTER

Color and flavoring problems in the food industries. T. H. FAIRBROTHER. *Food Manuf.* 6, 72-5(1931) E. H.

Possibility of use in foods of coloring materials made in Italy. C. MANUELLI AND R. MAGGIORA. *Atti III congresso naz. chim. para officinali* 1930, 627-35.—The characteristics of coloring materials allowable in foods are reviewed. E. M. SYMONS

Reaction between metal receptacles and foodstuffs. I. Relation of metals to fluids (drinks): corrosion. B. BLEYER AND J. SCHWABOLD. *Biochem. Z.* 230, 136-45 (1931).—The corrosion of various metals (Al, Zn, Fe, Ni, Sn, Pb, Cu, Ag, brass) by water, tea and coffee has been studied and the extensive findings are recorded in tables and curves. S. MORIGUTI

Determining the swelling rate in canned foods. W. S. STURGES, E. T. DRAKE AND L. B. PARSONS. *Food Ind.* 3, 105(1931).—A quant. method for the measurement of swelling rate of Sn cans is given. The principle of the method consists in weighing the can under H₂O and also in air. $S = 100 W/V$, where W is the decrement in wt. in H₂O, V = wt. in air - wt. in H₂O, and S is the % swell. C. R. FELLERS

The f_H values of New Zealand wheats and flours. J. E. DIXON. *New Zealand J. Sci. Tech.* 12, 146-53(1930).—The f_H and buffer-effect values of ground whole-wheat flour prepd. in the lab. and com. flour were detd. Exts. prepd. by the Kent-Jones method were used in a quinhydrone cell coupled with a calomel half-cell. One hr. was required to get a correct reading. The Pt electrodes poisoned readily. The results showed that the quality of a com. flour as detd. by the baking test cannot be predicted from the f_H or buffer-effect value of its wheat or flour prepd. in the lab. The f_H and buffer-effect values of flours are governed by the treatment each wheat receives at the mill and bear no relation to the values found for samples prepd. in the lab., although the lab. values may be of some assistance to the miller in choosing the method of treatment to be adopted for each lot of wheat. L. B. GILSON

Utilization of the soy bean. IV. Electrolytic oxidation of the hydrolyzate of proteins. I. YOSHITARO TAKAYAMA. *J. Soc. Chem. Ind., Japan* 34, Suppl. binding, 31-2(1931); cf. C. A. 24, 2810.—Crude soy-bean protein was hydrolyzed with H₂SO₄. The hydrolyzate was filtered from human and electrolyzed with a peroxidized Pb anode and a Pb cathode in an undivided cell (anodic c. d. 2 amp./sq. dm.), the soln. was then steam-distd. The distillate was neutralized with NaOH and redistd. The volatile acids remaining as Na salts were 9.6% of the crude protein. A mixt. of free acids was obtained from the salts by addn. of phosphoric acid and fractionated at 95-195°, sepg. formic and isovaleric acids and AcOH, with small quantities of propionic and butyric acids. The residual soln. of the steam distn. was distd. over CaO under reduced pressure into HCl. The volatile base obtained was pure NH₄Cl. The relation between NH₄N

produced and elec quantity consumed was detd. In the same way, formic and isovaleric acids and AcOH were obtained from gluten hydrolyzate. **E. M. SYMMES**

"Vaingan" rice cultivation in south Konkarn. **R. M. HARGREKATTI** *Poona Agr Coll Mag* 22, 140-59(1930)—The term "vaingan" is given to rice grown by means of lift irrigation during the hot summer weather. As compared with rainy season rice, "vaingan" rice kernels and bran are richer in albuminoids and digestible carbohydrates and poorer in woody fiber and ash, particularly siliceous ash. "Vaingan" rice straw contains more digestible carbohydrates and albuminoids than rice straw from a rainy season. **K. D. JACON**

The determination of starch in cereal products. **C. W. HEED AND D. W. KENT-JONES** *J Soc Chem Ind* 50, 15-22T(1931)—Methods for detg starch are broadly classified as (I) hydrolytic methods, (a) acid hydrolysis, (b) enzymic hydrolysis; and (II) non hydrolytic methods. The "by-difference" method and the method of direct washing out from dough are in the latter category. Some of these methods are discussed in detail and the relative results obtained by the application of them to different mill stocks are considered. These stocks included: (1) corn starches, (2) corn flours and mill stocks, (3) wheat offals, and (4) whole wheats. Conclusions. By "starch" different observers mean different things. The modified Rask method is suggested as of the greatest utility, while hydrolytic methods are open to grave suspicion. The modified method is: Place 1 g of material and 1 g acid washed sand in a centrifuge tube, mix and cover with washed Et_2O , stir, centrifuge and pour off the liquid. Make 2 addnl. washings and to the residue add 25 cc. of water and 0.25 cc. *N* NaOH, mix and after 15 min add 5 cc. pure MeOH and mix again. Mix in 5 cc. dil MeOH (2:1 water), centrifuge this mixt. and remove the alc layer. Wash the residue twice with 10 cc. of dil MeOH and finally three with water. Stir the residual mass into a thick paste with a few cc. of water, add a total of 20 cc. of water to transfer to a 100-cc. flask, add 20 cc. concd. HCl and make up the vol. to 100 cc. with Rask's acid. Filter the contents of the flask by suction through a Gooch crucible contg asbestos and a layer of acid-washed sand into a small dry suction flask. Transfer 50 cc. to a 200-cc. beaker contg 110-5 cc. 96% alc., and immediately after pipet drainage, stir the resulting mixt 1 min or until the ppt becomes flocculent. When it has settled, transfer the contents of the beaker to a centrifuge tube, centrifuge 10 min and remove the supernatant liquid. Wash the residue with 70% (by vol.) alc. and twice with 96% alc. Wash the ppt. with Et_2O , dry the crucible at 40° for 10-15 min. and then at 130° to const wt. **W. H. BOYNTON**

The determination of starch in flour by diastase-acid hydrolysis. **B. G. HARTMANN AND F. ILLIO** *J Assoc Official Agr Chem* 14, 112-6(1931)—A detailed description is given of a modification of the technic of the official A. O. A. C. method for the detn. of starch by diastase acid hydrolysis. The time required for starch conversion by diastase is reduced from 2 hrs. to 20 min. by introducing a preliminary pepsin digestion (cf. *C. A.* 21, 287), but the total working time required for the peptic-diastatic digestion is about 3 hrs. The peptic digestion increases the starch yield appreciably, this is shown to be due to the fact that the starch is rendered more accessible to diastatic hydrolysis and not to any reduction by the pepsin or the products of its action. In the pepsin digestion an acidity adjustment corresponding to 5 cc. *N* HCl per 100 cc. of substrate is made, after the digestion the acidity is carefully neutralized, the mixt is slightly acidified with HCl, and finally treated with CaCO_3 to assure a practically neutral condition for the diastatic inversion. The official method was also modified in regard to the removal from the flour of the sugars and free fat; the procedure proposed, though time-consuming, is easy to carry out and gives satisfactory results with ordinary care, even in the hands of analysts inexperienced with it. **A. PAPINBAU-COUTURE**

Starch explosion hazards reduced by safety measures. **D. J. PRICE AND H. R. BROWN** *Food Ind.* 3, 106-7(1931)—Explosions of dry starch are relatively frequent and are usually caused by sparks from machines, motors, etc., or by static electricity. A serious explosion at Decatur, Ill. is described. Safety measures are outlined. **C. R. FALLER**

Estimation of the age of flour. **W. HARTMANN** *Z. Untersuch Lebensm* 59, 304-79(1930)—Samples of flour and bread (33.3 g.) were extd. for 6 hrs. with boiling petr ether or CHCl_3 , and the *n* at 40° and the acid value of the ext. detd. Flours recently mill ground had acid values of 54 to 76 and 46 to 64 when contg 60% of rye and wheat flour, resp., while the corresponding figures for flours of the same compn., ground in the lab. and left for 24 hrs., were 44 and 39. Bleached flours were not included. Extras, after various intervals, showed that the acid value increases with time of storage, rapidly at first, to about 240 after 15 months. Gluten N, lecithin phosphate contents and

n decrease on storage at a rapid rate, particularly in the initial stages. The acid value and n of the unground grain, however, showed little alteration on storage. Reductions in n were observed for bread baked from rye and wheat flours, while the acid value was increased in the latter case only. The lecithin phosphate fell from 6 to 1.3 mg per 100 g of flour in both cases. The bacterial contents were also detd on samples 5 to 15 months old; in the presence of growths of *Oidia*, small quantities of an unsaponifiable wax-like substance were obtained. In view of statements that flour is no longer suitable for baking purposes after 6 months' storage, it is suggested that the max permissible acid values in such cases should normally be 120 to 150 for 60% rye flours, and 90 to 100 for 60% wheat flours. C. R. FELLERS

Direct determination of available carbon dioxide in baking powder. MAYNARD R. COE. *J. Assoc. Official Agr. Chem.* 14, 99-101 (1931).—In the gasometric method adopted by the A. O. A. C. for the detn of CO_2 , 2 detns must be carried out to obtain available CO_2 , which is the difference between the total and residual CO_2 . The following new method is proposed, the same app is used and only 1 detn is required. Run 25 cc. of 5% $(\text{NH}_4)_2\text{SO}_4$ soln into the 250-cc. decompn flask contg 17 g of baking powder, heat at boiling temp till all gas is evolved, cool to room temp; when equil is established (generally in about 15 min) read the amt of available CO_2 on the graduated tube and make the necessary temp and pressure corrections. Total CO_2 also can be detd in a few min with the same sample, without detaching the flask from the app, as follows. Fill the buret contg a few cc of $(\text{NH}_4)_2\text{SO}_4$ with H_2SO_4 (1 + 5), run 25 cc into the flask, heat almost to boiling, let cool as before, take the reading, subtract a no corresponding to the no of cc. of acid used from the no of cc. found, and correct for temp and pressure. $(\text{NH}_4)_2\text{SO}_4$ is used as reagent because (1) a satisfactory evolution of CO_2 is obtained from all varieties of baking powder; (2) it acts as a protein coagulant for powders contg egg albumin, thus serving as a foam destroyer, (3) if distd water is used with phosphate powders and those contg egg albumin, low results are obtained, but with 5% $(\text{NH}_4)_2\text{SO}_4$ soln the same results are obtained as with the gasometric method, (4) if the total CO_2 value is desired the acid reagent may be used to evolve the residual CO_2 remaining in an $(\text{NH}_4)_2\text{SO}_4$ soln. The technic of the method is described in detail. It gives results checking with those of the gasometric method as closely as would duplicates by the latter method. A. PAPINEAU-COUTURE

The function of emulsions of oil and water in breadmaking, with special reference to gluten formation and modification. J. C. VAN DYKE. *J. Soc. Chem. Ind.* 49, 421-2T (1930).—Glutens obtained from doughs which had been made up with emulsions are softer and more elastic. When the gluten is dried in a hot air oven at 105° , the color of the baked gluten is superior to that obtained by drying off the glutens from a dough of flour and tap water. It is very difficult to wash out doughs made from flour and emulsions, and in order to prevent dispersion the starch must be filtered through fine muslin. Emulsions of oil and water increase the amt of gluten in the flour to a very small extent. This contradicts Bension (*C. A.* 21, 613). There seems to be no definite relation between the fact that emulsions of oil and water keep bread moist for a long period, and the increase of gluten. As the washing free from starch is very difficult, it seems that the increase of gluten reported by B was caused by some const. error connected with the washing out of the gluten. ALBERT L. HENNE

Phosphatide content of different breads. B. RAWALD. *Z. Untersuch. Lebensm.* 60, 315-8 (1930).—The phosphatide contents were detd by R.'s method (*C. A.* 23, 3940), the sample being extd in succession with a large excess of acetone and a mixt. of C_2H_5 and EtOH . The exts were evapd separately, the residues extd with dry Et_2O , and the P was detd in each case in the residue after evapn. of the Et_2O . The sum of these 2 figures was taken as the total phosphatide content, the following percentage of the dry samples being found: Rye bread crumb or crust 0.174, crumb and crust of white bread 0.287, crumb and crust of pumpernickel 0.073, black bread crumb 0.183 and crust 0.194, graham bread crumb 0.267 and crust 0.302. The respective P contents of the residues after extn were 0.20, 0.24, 0.32, 0.34, 0.35, 0.30 and 0.29%. With the exception of pumpernickel, the different breads examd. contained approx. the same amts of phosphatide. Similarly, the P contents of the residues were approx. equal in all cases. Temp is without influence on the phosphatide content because both crust and crumb had the same amts. Yeast fermentation was also without effect on the phosphatide of bread. C. R. FELLERS

The extinction coefficient of milk. A. SCHNECK and H. MENGEBAIER. *Milchwirtschaft. Forsch.* 11, 1-29 (1930).—The light transmission is proportional to concn. under strictly controlled conditions. The Lambert-Beersche rule may be employed. The fat phase has little effect. In dilns of 1-100 the results satisfy the Lambert-Beersche rule. Many

samples tested show the extinction to be depressed 4% by skimming. Milks from different cows have different extinction coeffs. It is shown that the extinction coeff due to fat varies with the av. size of the fat particle. GEORGE R. GAERNBANK

The residual nitrogen of cow milk. F. KIRPPELE AND J. GLOETZL. *Milchwirtschaft Forsch.* 11, 62-117(1930).—The authors consider the residual N of milks as analogous to that of blood and urine. They ppt. the coagulable protein by heat and 10% AcOH. The N remaining in the filtrate is called total residual N. The total residual N for an individual milk is 84.8 mg % per 100 cc when the total N is 531.7 mg % for herd milk it is 60.3 mg % for a total N of 533.1 mg %. If the hexone bases are removed by Na tungstate and trichloroacetic acid the residual N for an individual milk is 28.4 mg. %. The values for the total residual N in mg % of the constituents for individual and herd milk are, resp., albumoses 30.7, 27.2, peptone 24.2, 18.7, amino acids 3.0, 4.6, NH₃ 1.3, 1.1; urea 13.4, 10.1; creatinine 2.3, 1.7; creatine 3.1, 2.2, uric acid 2.1, 1.5. The total residual N increase in pasteurization is 10.4% and in boiling 18.6%. GEORGE R. GAERNBANK

Proteolysis of milk. H. HAWKICK. *Milchwirtschaft Forsch.* 9, 339-51(1930).—The detn. of caseolytic organisms in milk is described; the detn. of NH₃ by distn. in a vacuum and the detection of peptone are also of value. Tryptophan was not detected. The catalase value is not always parallel with the no. of proteolytes.

B. C. A.

Has "chromiform," the new milk preservative, proven its value? FONZES DIACON. *Ann. fals.* 24, 22-5(1931); cf. C. A. 23, 913.—The previous investigation established the value of "chromiform" when added to fresh milk. Evidence is now presented showing that, when added to milk several hrs. old, "chromiform" is a much better preservative than either K₂Cr₂O₇ or (ClH₃O)₂ alone. A. PAPINEAU COUTURE

Influence of formaldehyde on the precipitation of milk protein material. MASCHL MASCHL AND EMILIE BOUCHIARA. *Bull. soc. chim. biol.* 12, 994-1000(1930); cf. C. A. 24, 5055.—Addn. of up to 20% formal to milk caused more AcOH to be required for pptn. of casein, but less N was left in soln. after pptn. was complete, and there was less re-soln. upon adding an excess of acid. A similar effect was exerted upon the pptn. of total milk proteins by CCl₃COOH. C. G. KING

Souring of milk. I. W. GEMMER AND W. PAAP. II. W. GRIMMES AND C. ASLART. *Milchwirtschaft Forsch.* 9, 88-99, 100-20(1930); cf. C. A. 23, 3990, 24, 3025.—Rennin coagulum and the residue of milk on porous plates differ in that the N, Ca and P contents of the latter are the smaller. Possibly in coagulation by rennin dissolved Ca and P are co-pptd., or combination of Ca₃(PO₄)₂ with caseinogen or casein occurs. With progressive souring the Ca content of the coagulum diminishes fairly regularly. If $\sqrt{K_S}$ is the H ion concn. of lactic acid of acidity α , c is a factor and y is the H ion concn. of the milk, $y = \alpha x^2 \sqrt{K_S}$. The value of c was 3.03-5.42 (av. 4.3) $\times 10^{-4}$. B. C. A.

The foaming problem. W. MOHR AND C. BROCKMAN. *Milchwirtschaft Forsch.* 11, 48-61(1930).—Milk with the greatest surface tension and viscosity produces the most stable foam. The foaming properties are greatly changed by the addn. of Na citrate. The solidity and stability of the foam are greatly diminished by the salt. There is a relation between the foaming and the surface viscosity. The tendency to foam depends on the ability to form a surface. Solidification of the wall gives a measure of the surface viscosity but not of the stability. It requires an optimum surface viscosity for a stable foam. With increasing fat foam formation and stability decrease. Foam formation is greater when the fat is liquid than when solid. Homogenization of skim milk reduces foaming, with whole milk it is increased. The influence of temp. varies with different milks. The optimum is from 30° to 60°. Repeated foaming does not decrease the capacity. Churning reduces the capacity of the skim milk. G. R. G.

Detection of heated milk. K. EBLE AND H. PREIFFER. *Z. Unters. Lebensm.* 60, 311-4(1930).—Of the several variations of the benzidine test for the detection of heated milk, the following method gave the most satisfactory results: shake 5 cc. of milk with 0.5 cc. of a 4% soln. of benzidine in 96% EtOH, add 1 drop of a 1% soln. of H₂O₂ after 1 min., and shake the mixt. again. Raw milk gives a stable blue-green color, milk heated for 30 min. at 63° gives a transient blue-green color, while strongly heated milk is colorless. After 1 min. add 10 cc. of a satd. soln. of MgSO₄ and shake the mixt. raw milk becomes gray green gradually turning to blue green in about 10 min.; heated milk is gray green and gradually becomes grayish blue. After 10 min. more dark blue curds sep. from the slightly heated milk, while in the strongly heated milk the curds are white. If MgSO₄ is replaced by Na₂SO₄, the respective colors are chocolate

to violet-brown and green-brown. Heated milk having an abnormally high leucocyte content or lactic acid acidity, or milk heated for a short time below 63°, tends to give the color reactions of raw milks, and it is desirable before making the test to add 0.25 N KOH soln. to reduce the acidity to 0.06% lactic acid. These exceptions are due to the peroxidase assoc. with the proteins or leucocytes of the milk, which normally passes into soln. when the milk is heated and hence loses its activity, but may be activated by lactic acid and inactivated by the addn. of AmOH. In doubtful cases a mixt. of 10 cc. of milk, 10 cc. of a satd. soln. of $MgSO_4$ and 1 cc. of a 20% soln. of AcOH is shaken and centrifuged for 6 min. at 120 r. p. m. Slightly heated milk gives a clear serum, raw milk a definitely opalescent serum, and mixed milk contg. 20% or more of raw milk, a faint turbidity.

Pseudomonas as a cause of bitter milk. R. W. NEWMAN. *Calif. Dept. Agr. Monthly Bull.* 19, 640 (1930).

Anaerobic spore test as an index of contamination in milk. M. H. KNUTSEN AND E. C. HOLST. *Penn. Agr. Expt. Sta., 43rd Ann. Rept. of the Director in Bull.* 258, 29 (1930).—The addn. of reduced Fe increased the efficiency of the test, but even this improved technic failed to show any decided degree of accuracy in differentiating milks as to the amt. of manural contamination. One of the principal reasons the test is invalid is because anaerobic spores germinate after varying periods of time depending upon the temp. of incubation and other factors.

The physicochemical constitution of spray-dried milk powder. L. H. LAMPITT AND J. H. BUSHILL. *J. Soc. Chem. Ind.* 50, 45-54T (1931).—It is shown that the figures obtained for fat content (Soxhlet) are not comparable to the "free" fat as detd. by the standard method as described. The normal moisture content of com. milk powders has no effect on the amt. of "free" fat. The "free" fat content of 8 com. spray-process powders is given, showing that the percentage calcd. on the total fat may vary between 3.33 and 14.22, for roller process powder it is between 91.6 and 95.8. Increasing the surface of the particles of spray-process powder by grinding 3 hrs. increases the "free" fat up to 83% of the total, but in the roller process powder practically all of it is free. The differences in the systems of drying which may influence the constitution of the particles are pointed out. Attempts to simulate the Soxhlet mechanism by a repetition series of standing extns. has thrown some light on the differences observed. It is shown that when milk powder absorbs moisture certain well defined changes take place and that a first "clummy" stage is succeeded by the development of a dry, hard and powdery texture. This change is accompanied by a freeing of the fat. It is proved that there is a definite range of moisture content at which the fat is freed; this is called the "critical moisture content" by the authors. Microscopical examn. proves this change to be accompanied by the crystn. of the lactose, and this crystn. is also produced by treatment with 96% alc. Spray process lactose is amorphous but crystallizes under the effect of exposure to moist air or to treatment by 96% alc. Time is a factor in the freeing action of absorbed moisture; powder freshly made reacts more slowly than an older one. The fat-freeing action of moisture is irreversible. The crit. moisture range is independent of the fat content and proportional to the amt. of solids not fat in the powder. It is not possible to put forward a definite theory concerning the structure of the milk-powder particle. The lactose in its amorphous state has an important effect on the availability of the fat to solvents, and any means so far tested by which the lactose is caused to crystallize results in the freeing of the fat.

J. C. JURRIENS.

Butter-fat test in condensed and evaporated milk. W. D. SWOFF, C. D. DAHLE AND F. J. DOAN. *Penn. Agr. Expt. Sta., 43rd Ann. Rept. of the Director in Bull.* 258, 26 (1930).—The method for sweetened condensed milk is that previously given (C. A. 24, 4869). In testing evapd. milk, excellent results were obtained by using 3 cc. of NH_4OH , 4 cc. of BuOH and 17.5 cc. of 1.1 H_2SO_4 added to 9 g. of evapd. milk. The fat column was measured, as for condensed milk, from the bottom of the lower to the bottom of the upper meniscus. This method also yields accurate results when used on ice cream and ice cream mixes.

C. R. FELLERS.

Sterilization of glass-lined tanks. F. M. SCALES AND HARRIET E. RUSSELL. *Food Ind.* 3, 120-1 (1931).—Results indicate that glass-lined tanks which have been properly cleaned and rinsed and then sprayed with a Cl soln. contg. 223-250 p. p. m. of available Cl in an alk. NaOCl soln. will be sterile or nearly so 30 min. after being sprayed. Bacterial cultures obtained from non-sterile milk bottles were used in carrying on the expts.

C. R. FELLERS.

Advances in processing aid domestic casein. F. L. CHAPPELL. *Food Ind.* 3, 102-4 (1931).—Approx. 60,000,000 lb. of casein, 80% of which is consumed by the paper

trades, is annually consumed in the U S. A casein low in both ash and acidity is required. A continuous process of manuf. from skim milk is described. Briefly, the casein is pptd by HCl , the concn being approx 0.5%, at a temp. of 43°. The curd is broken up, washed, pressed and dried to 10-15% moisture content within a very few min. The quality of the domestic product is constantly being improved.
C. R. FELLERS

Bacterial flora of some ice cream ingredients. R. W. NEWMAN AND A. E. REYNOLDS. Calif Dept Agr, *Monthly Bull* 19, 677-83(1930)—Data are given for frozen egg yolks, fresh eggs, powd eggs, powd skim milk, gelatin, sugar, powd. cocoa, cold pack (frozen) strawberries and ice cream colors.
C. R. FELLERS

Gelatin percentage can be reduced. E. O. ANDERSON, T. A. LYONS AND R. L. PIERCE. *Ice Cream Trade J* 26, No 11, 81-3(1930)—Ice cream of good body was secured with the grade of gelatin used when the percentage of gelatin was about 0.10% and the homogenizing pressure was 3000 lb. Apparently satisfactory ice cream contg no gelatin was also obtained when the homogenizing pressure was 3500 and 4000 lb. By taking advantage of these data com ice cream is now being prepd in which the gelatin has been reduced to 0.10%.
A. H. JOHNSON

Gelatin will not be replaced. B. E. MORRALL. *Ice Cream Trade J*, 26, No. 11, 80-1(1930)—Expts were conducted in which it was attempted to replace gelatin in the ice cream mix by using high homogenization pressures. The results indicated that the lots with higher gelatin contents and lower homogenizing pressures were not so coarse or icy as lots with lower gelatin contents and higher homogenizing pressures. The smoothness in texture of the ice cream increased with the quantity of gelatin used. Water showed a tendency to drain out of mixes contg the smaller quantities of gelatin. The lower the gelatin content of the mix the shorter was the time required to reach the desired overrun. In general, the work indicated that high homogenizing pressure would not entirely replace gelatin.
A. H. JOHNSON

Consistency of Swedish butter. E. HAGLUND, G. WODE AND T. OLSSON. *Kgl. Landbruks-Akad Handl Tid* 69, 1147-72(1930)—The consistency of Swedish butter at different times of the year and from various districts, and factors causing these differences, were detd. The no. of samples was 413. Butter hardness was detd by the Perkins method (C. A. 8, 1313) at 15°, av. of 6 impressions, expressed as the wt in g. which the falling body must have for the rod to displace 1 cc. of butter. The I value of the butter was detd also. Butter hardness varies considerably with the time of year, and that from different districts at the same time of year is different. These variations are caused by variations in value of the fat and hence are due to feeding. Hardness decreases as I value increases. Fat with a given I value can give butter with varying degrees of hardness. Churning tests showed that permanent variations in butter hardness can be produced by manuf. methods. Long cooling of cream before churning increases butter hardness, as does long cooling of butter granules before working. Differences in degrees of butter hardness caused by methods of prepn are as great as variations in hardness of collected samples of identical I value.
E. M. SYMMES

A new relation for the determination of vegetable oils in butter. OSCAR DE SOUZA VIEIRA. *Rev. brasil. chim* 2, 237-46(1930)—Taking into consideration the increasing use of babassú oil and other palmaceous oils in the adulteration of butter and paying particular attention to the relative meanings of the various indexes, S. V. studies the methods actually used for the discovery of such fraudulent mixts. Bromometry is not yet suited to the study of small quantities of vegetable oils. A valuable relation results from the great differences between the consts. of the insol. fatty acids (Polenske) and butyric and caproic acids (xylene) in various kinds of vegetable oils, principally palmaceous oils. Thus the relation (Polenske/xylene) $\times 100$ is found to be 10 for the various kinds of butters and possibly 1529 for babassú oil.
JOHN M. LADRO

Effect of method of preparation on the calcium content and consistency of cheese. GUNNAR WODE. *Kgl. Landbruks-Akad Handl Tid* 69, 1032-43, *Medd. Centralanstalt. försöksväsendet jordbruks. Mejeriförsök* No. 40, 14 pp (1930)—There is a close relation between Ca content and body of cheese, the higher the Ca content the weaker the body, and vice versa. During cheese making the vat whey dissolves Ca from the curd. A loss of Ca is caused also by the whey removed during the pressing. The quantity of Ca dissolved from the curd by the vat whey increases, parallel to the acidity of the milk before setting. If the vat whey is sepd quickly from the curd it dissolves less Ca. The quantity of Ca dissolved from cheese by press whey increases with increasing vol. and acidity of press whey. The total Ca dissolved by vat whey is considerably greater than that dissolved by the press whey.
E. M. SYMMES

Bleaching, muddy discoloration and black-spot development in colored cheese.

G F V MORGAN *New Zealand J Agr* 42, 35-6(1931), cf *C A* 25, 748—Bleaching of annatto-colored cheese was traced to the reducing action of certain organisms, including *B subtilis* and *B fluorescens liquefaciens*, on the dye in the absence of an adequate supply of atm O. Upon longer keeping the bleached areas acquired a muddy discoloration. Black discoloration was traced to the action of bacteria in reducing Pb salts present in the annatto coloring material. K D JACOB

Physicochemical investigations of vegetable juice. E. CANALS, J CANAYE AND I CADANES *Bull soc chim biol* 12, 1022-4(1930)—P, Ca, Mg, Na and K were detd in the expressed juice and in the dialyzate from turnips, carrots, radishes, spinach, asparagus, peas, beets and 6 fruits. Practically all of the Na and K and over 90% of the P and Mg were removed by dialysis. Of the Ca 30-94% (av 70) was removed. C G KING

Maraschino cherries, their preparation and manufacture. E H WIEGAND AND D E BULLIS *Glass Packer* 4, 125-7, 134(1931), cf *C A* 24, 439—For bleaching the raw unstemmed cherries (white sweet varieties) a soln of 1.5% of SO_2 and 1.5% soln of $\text{Ca}(\text{OH})_2$ is used. The process is usually carried out in hbls. Excessive lime causes spoilage to occur from gas forming and liquefying bacteria. Daily agitation of the hbls is necessary to prevent the lime from settling out. The cherries keep practically indefinitely in this soln. The cherries are leached in cold water for about 20 hrs to remove SO_2 and lime, stemmed and pitted, and finally dyed with erythrosin or mixts of erythrosin and Ponceau in a slightly alk soln. NaHCO_3 added at the rate of 7 oz per hbl is used to give proper reaction for dyeing. About 1 oz of erythrosin per 100 lb of liquid is used. The dye is set by acidifying with citric acid added at the rate of 1 lb per 100 lb of liquid. The cherries are now ready for the candying process, which consists of the gradual building up of a sirup concn of 50-65% of sucrose. Oil of bitter almonds or Bzl is universally used for flavoring purposes. Sterilization of the bottled product is carried out at a temp of 180° for 25-30 min. C R FELLERS

Determination of formic acid in fruit juices. A HANAK AND K KIRSCHNER. *Z. Untersuch. Lebensm.* 60, 278-90(1930), cf *C A* 24, 4098—Steam-distill with 0.3 g of tartaric acid for 45 min. 10 cc of the juice or 10 g of the fruit sirup dild to 20 cc; the increase in the vol in the distn flask should not exceed 20 cc. Avoid overheating by the use of an asbestos shield and collect 400 cc. of distillate in just enough distd. water to cover the end of the adapter. Dil the distillate to 500 cc., titrate 200 cc. of it and calc. the total free volatile acids. To 100 or 200 cc. of the distillate, according to the amt. of HCO_2H present, add the calcd amt of 0.5 N NaOH soln required to neutralize the acidity, followed by 0.5-1.0 cc in excess. Warm the mixt for 10 min. to saponify any formic esters and then evap. with 0.2 g. of Na_2CO_3 without boiling, until 30 cc. remains. Add sufficient 0.2 N KMnO_4 (usually 3-5 cc.) so that an excess of 1-1.5 cc. is present. Oxidation is complete after 45-60 min., then add 1 cc. of 10% ZnSO_4 soln. to flocculate any suspended matter, and dil the soln. to 50 cc. and filter. Det. the excess of KMnO_4 colorimetrically by matching the color of 25 cc. of filtrate against that of a 0.2 N soln. of KMnO_4 in a colorimeter. The colorimetric method is sensitive to 0.1 cc. (equiv. to 0.276 mg of HCO_2H). The method gave accurate results on a variety of fruit juices. C R FELLERS

The manufacture of lime juice and its economic value. V H KULKARNI *Poona Agr. Coll. Mag* 22, 88-92(1930)—The manuf. and preservation of lime juice in India are described. The yield of clear juice is about 40% of the wt of the fruit. West Indian bottled lime juice contains about 5.6% citric acid while the av. Kagdi lime juice contains approx 7.5%. K D JACOB

New method of enzymic clarification of unfermented apple juice. ZOLTAN I. KERTESZ. N. Y. State Agr. Expt., *Bull* 539, 3-10(1930)—A very active pectin-decomp. enzyme obtained from *Penicillium glaucum* is proposed as an agent for assisting in the clarification of apple cider. The enzyme can be prepd. from the mold in either solid or liquid form. The method of extn is not given. Starches or proteins are not attacked by the enzyme prepn in concns of 0.5% acting at 55° for 10 hrs. The insol. matter which is formed by the decompn. of the pectin and is largely responsible for the cloudiness in cider is readily removed by filtration or centrifugation. The clear filtrate may be bottled and pasteurized for 20 min at 76° without showing any change in flavor. In 1 test about 30% of the pectin was removed from cider by enzyme clarification. The method may have com. applications. C. R. FELLERS

Removal of spray residues from apples; a wax-solvent method. J. R. NELLER. *Ind. Eng. Chem* 23, 323-5(1931)—Winesap apples were dipped in various solvents to remove the exuded wax and the oil that had been added in the spray mixts. MeOH proved to be a desirable solvent and enabled the HCl wash effectively to remove arseni-

cal spray residues that could not be removed with HCl alone even when heated. The storage quality of the fruit was not impaired. Practical considerations of the method are discussed. J. R. NELLER

Composition of tomatoes and tomato products and the determination of water in these products. A. LEONHARD. Z. *Untersuch. Lebensmittel.* 60, 185-93(1930)—The compn. of 34 samples of German and Italian tomato products is tabulated. After comparing the various methods of detg. dry solids in tomato products, L. prefers to det. the water directly by distn. with xylene. Place 5-10 g. of the well mixed sample in an Al boat and dist. with 150 cc. of xylene in a 300-cc. Erlenmeyer flask on a sand bath. A delivery tube 350 mm. long, bent at an angle of 45°, leads into a vertical 150-cc. graduated receiver completely immersed in running water and closed except for a glass tube drawn out at 1 end to 2 mm. in diam. and inserted through a rubber stopper. The bottom of the receiver contains 1 cc. of 1:1 KOH soln. and 150 cc. should distil over in about 20 min. Remove any condensate from the delivery tube by application of a small flame, shake the contents of the receiver gently and read the vol. of the sep'd aq. layer after 2 hrs., deducting the 1 cc. of alkali originally present. Volatile oils are retained in the xylene. The results for dry matter were compared with those obtained by calcn. from the sp. gr. of the filtered aq. ext. In general, the dry solids obtained by the direct distn. method were 0.0-4.7% less than by the sp. gr. method, i. e., the moisture results were probably a little high, especially for samples of the more liquid products. C. R. FELLERS

Converting wastes into profits. A study of what can be done with tomato pomace. EARL D. STEWART. *Food Ind.* 3, 112-4(1931)—Tomato seeds contain approx. 18-20% of high-quality edible oil resembling cottonseed oil, solidifying pt. -9°. It contains about 45% olein and 34.2% linolein. Tomato pomace may be readily dried in a rotary air drier. The yield is 15 lb. per ton of tomatoes. The compn. is crude protein 21.0, crude fat 18.0, fiber 27.3, N free eat. 24.9 and moisture 50%. The product is ground in a hammer mill. Although the taste is bitter, the product is suitable for cattle feed. Other products are considered briefly. C. R. FELLERS

Insect control in dried fruits. H. M. REED. *Food Ind.* 3, 117-9(1931)—The use of cold storage, vacuum fumigation with CS_2 and CO_2 , HCN, ethylene oxide and ethylene oxide- CCl_4 mixts. are advocated for preventing or destroying insect infestation in various dried fruits under varying conditions. C. R. FELLERS

Some properties of honey colloids and the removal of colloids from honey with bentonite. R. E. LOTIKOFF AND H. S. PAINE. *Ind. Eng. Chem.* 23, 328-32(1931)—Honey samples contain 0.1-1% colloids, which exert a considerable influence on the properties of the honey. The particles have an isoelec. point at pH 4.3, at which max. flocculation by diatomaceous earth and min. cataphoresis occur. Addn. of bentonite flocculates the colloids completely. K. V. TRIMANY

Some organic acids in honey. E. K. NELSON AND H. H. MOTTERN. *Ind. Eng. Chem.* 23, 335-6(1931)—The non volatile acids were detd. by Hartmann and Hillig's method (C. A. 24, 2208), the pptn. with $Pb(OAc)_2$ being checked with known mixts. of malic and citric acids in the presence of sucrose. Succinic acid is present in the more acid specimens. Volatile acids, mostly $HIOAc$ and HCO_2H , range from 0.01 to 0.05%, malic and citric up to 0.05 and 0.008%, resp. K. V. TRIMANY

Ultra-violet absorption of honeys. J. STITE AND J. KOZEK. Z. *Untersuch. Lebensmittel.* 60, 420-5(1930)—The Bunsen-Rosecoe absorption extinction coeff. ϵ is given in terms of the intensities of incident light I_0 and transmitted light I , the concn. c , and the thickness of the layer of soln. exam'd x by $(1/cx)\log(I_0/I)$. By means of the Judd-Lewis sector photometer and Hilger quartz spectrograph ϵ was detd. for wave lengths λ of 5500 to 2350 Å. U. from the measured values of $\log(I_0/I)$ with a relative error of $\pm 2.2\%$. A layer 1-4 cm. thick was used, the source of illumination being a W. Fe arc with electrodes 3 mm. apart in front of a 0.02 mm. slit. Absorption curves in which ϵ is plotted as ordinate against λ , are shown for 6 varieties of genuine honey, a characteristic curve always being obtained with a sharp max. at 2700 Å. U. corresponding with $\epsilon = 0.13$ to 0.145 for a 1-cm. layer of a 4.5-6.5% soln. Glucose, fructose and sucrose give higher absorption maxima than water but all are relatively low. For com. sucrose, ϵ is slightly higher than in the pure product. Dextrin has a very high absorption at 2700-800 Å. U. Addn. to honey of 22% of sucrose lowered ϵ by 0.3. The total absorption of honey cannot be detd. from the individual absorptions of its sugar constituents alone, since the nature and concn. of the ash, coloring matter and proteins exert a considerable influence. C. R. FELLERS

Mushrooms should be packed in glass. C. G. HICKS. *Glass Container* 10, No. 1 8-9, 22(1930)—Darkening or oxidation of canned mushrooms is due solely to the

presence of O and not to the action of light. A weak citric acid brine is used in the prepn. to prevent darkening during the heat treatment, which varies from 30 to 45 min. at 240° F. C. R. FELLERS

Stem marrow kale fodder, its cultivation and yield. H. EDIN and G. SUNDELIN *Kgl. Landbruks-Akad. Handl. Tid.* 69, 972-1013(1930).—Stem marrow kale has been used widely recently, but data on feed value, cultivation and yield are scarce and inadequate. The dry substance of fresh stem marrow kale is 11.5-15%. Leaves contain twice as much protein as the stalk, and considerable fat. The org. substance was crude protein 18.4, fat 3.1, fiber 18.5, N-free ext. 62.0 and true protein 10.9%. The org. substance was about 86% of the dry substance. Digestibility coeffs. of late harvested kale were crude protein 75, fat 66, fiber 48, N-free ext. 88, carbohydrates 81, org. matter 78 and true protein 67. Fresh stem marrow kale contains about 95 feed units per 100 kg. of org. matter, 7.5 kg. of digestible protein, and 68 starch equivs., whereas dried kale has 66 feed units. The numerical value is 87-91 for fresh kale and 76 for dried kale. For cattle feeding, grinding or cutting is unnecessary, but the milk may have the taste of kale. Sheep eat stem marrow kale, but coarse stalks must be crushed. Pigs eat leaves without grinding, then eat stalks later. For 50 kg. body wt., 5 kg. of kale per day may be fed pigs, but this fodder is not a fattener. Poultry will eat cut kale, as a supplementary, rich vitamin food. Detailed instructions for cultivation of stem marrow kale are given. E. M. STRAINES

The camelthorn (*Acacia gerrardii*, Burch), J. S. HENKEL. *Rhodesia Agr. J.* 28, 71-3(1931).—The pods of camelthorn are eagerly eaten by elephants, giraffes and other wild animals in Africa. Their compn. is: H₂O 9.4, acid-sol. ash 3.3, ether ext. 1.6, fiber 31.0, crude protein 11.4, CaO 0.9, P₂O₅ 0.24 and K₂O 1.3%. A. L. MEHRING

Experiments in spray residue removal (STREETER, *et al.*) 15. A study of arsenical residue on apples in Pennsylvania (HODGKISS, *et al.*) 15. Importance of microbiology [in study of milk and its products] (CAN) 15. Life and work [on ice cream] of the late Professor A. C. Baer (BECKER) 2. Proteins of Indian foodstuffs (NARAYANA) 11D. Influence of neutral red on respiration [of bakers' yeast] (GEIGER-HUBER) 11D. Rust resistance in wheat (HANNA) 11D. The antagonistic substances formed during bacterial fermentation (KLEIN) 11C. Apparatus for pasteurizing milk or other liquids by "holding process" (U. S. pat. 1,792,328) 1. Vacuum tank and heated tube system for condensing milk or other liquid products (Brit. pat. 335,964) 1. Continuous process for sterilizing milk or other liquids under pressure (Brit. pat. 337,027) 1.

Concentrated milk industries (*New journals*). Published monthly by the Olsen Publishing Co., 501-15 Cherry St., Milwaukee, Wis. Vol. 1, No. 1, appeared in Sept., 1930.

BEYTHIEN, A.: *Laboratoriumsbuch für den Nahrungsmittelchemiker*. Dresden. T. Steinkopff. 589 pp. M. 33; bound, M. 40.

"Cerec"—Proceedings of the First International Conference on Flour and Bread Manufacture, in Prague, September 22, 1927. Edited by FRANCIS HRAUSKA. Prague: "Cerec." 578 pp. Reviewed in *Cereal Chem.* 7, 193(1930).

Index to the Literature of Food Investigation. Vol. II, No. 2. Compiled by AGNES E. GLENNIE. London: H. M. Stationery Office. 89 pp. 2s. Reviewed in *Food Manuf.* 6, 84(1931).

LABBÉ, HENRI: *Le cacao et le chocolat au point de vue alimentaire et hygiénique*. Brussels (19, Rue des Chartreux): Secrétariat général du Congrès international des fabricants de chocolat et de cacao. 24 pp.

OLIVAN, NICASIO: *Industrias de la fecha*. Barcelona. Cervantes. 204 pp.

Preserving food. FRITZ WALTER RAUTH and WILHELM SCHEFFER. Fr. 694,527, April 25, 1930. Packing materials for food are coated with a mixt. of glycerol, a colloid such as agar-agar and water.

Preserving foods. EATON FREUND. Fr. 695,096, May 5, 1930. Foods are preserved by heating them in a soln. of agents, such as salts of open-chain hydroxy acids, which facilitate the swelling of the dried products. Examples are given of the use of Na lactate and gluconate.

Conserving foodstuffs. RICHARD WILLSTÄTTER. Ger. 513, 656, Jan. 22, 1929. Animal and vegetable foodstuff is conserved by addn. of not less than 0.004% vol. of HCN in closed vessels. Thus, meat is kept fresh for 6 weeks at 20° with 120 to 300 mg. HCN to 5 kg. of meat in a 100 l. vessel.

Vitamin food. LAURENT M. RAYBAUD. Fr. 691,443, July 24, 1929. The vitamins in germinated seeds, whole or ground, are preserved by a coating of sugar chocolate, etc.

Hydrolyzed protein products. CHARLES N. FREY (to Standard Brands, Inc.) Can 308,910, Feb. 24, 1931. Hydrolyzed protein substances high in assimilable N are made by hydrolyzing a protein-contg substance with H_2SO_4 , neutralizing a portion of the H_2SO_4 with CaO , sepg the ppt and neutralizing a further portion of the acid with aq NH_3 .

Flour. METALLGES A-G. Fr. 691,666, April 29, 1930. Flour for cooking purposes is improved by adding to the flour or to the cereal before grinding, perphosphates up to 2.5% or more of the wt of the flour, so that besides O for the bleaching, there remains in the phosphate an amt of O sufficient to have a raising action on the flour during baking.

Dough-bleaching agents comprising bean flour, etc. D. VEAUV. Brit. 335,943, July 3, 1929. A dough bleaching agent is prepd by treating fresh untreated legumes contg carotin removing enzymes with a green malt infusion, drying the treated material at a low temp and grinding. NH_3 or Ca phosphates (3%) may be added to the infusion before use or similar mineral salts may be added to the bean flour in larger proportion, to form a composite yeast food and bread improver. Cf. C. A. 24, 1676.

Bread. OTTO GRACL. Ger. 516,268, July 27, 1928. In the manuf. of bread of relatively low carbohydrate content by baking a mixt of rye bran or groats with fresh wheat gluten, sterilized white of egg is added to the leaven.

Yeast product. CHARLES N. FREY, ELMER B. BROWN and COLGATE CRAIG (to Standard Brands, Inc.) Can 308,909, Feb. 24, 1931. A food product is prepd by heating a mixt of 4 lbs of yeast, 35 cc of HCl (d 1.19), 10 cc of lactic acid (85%) and 2.5 cc of phosphoric acid (85%), hydrolyzing the mass, sepg the residue and evapg the ext., neutralizing the excess of acid and admixing flavoring ingredients prior to the completion of the evapn.

Yeast product. CHARLES N. FREY, ELMER B. BROWN and COLGATE CRAIG (to Standard Brands, Inc.) Can 308,909, Feb. 24, 1931. A food product is prepd by heating a mixt of 4 lbs of yeast, 35 cc of HCl (d 1.19) and 23 cc of lactic acid (85%), hydrolyzing the mass, sepg the residue and evapg the ext., the excess of the acids being neutralized prior to the completion of the evapn.

Yeast product. CHARLES N. FREY, ELMER B. BROWN and COLGATE CRAIG (to Standard Brands, Inc.) Can 308,911, Feb. 24, 1931. A food product is prepd by hydrolyzing 4 lbs of compressed yeast with 40 cc of HCl (d 1.19) and 10 cc of H_3PO_4 (85%) in an autoclave at 15-20 lb pressure for 3 hrs. The ext. is sepd and the residue evapd to a pasty consistency of d 1.4. The acidity may be neutralized with soda, a favorable final reaction is approx neutrality to litmus.

Baking powder. WILLIAM E. STOKES and REGINALD A. WRIGHT (to Royal Baking Powder Co.) Can 308,902, Feb. 24, 1931. A baking powder comprises NaHCO_3 300, $\text{Ca}(\text{H}_2\text{PO}_4)_2$ 33, Na li pyrophosphate 405 and starch 262 parts.

Testing milk. N. GERBER'S CO., G. M. B. II. Ger. 516,423, April 15, 1928. To det whether udder disease is affecting the compn of the milk, samples from each of the four quarters of the udder are treated with an indicator, preferably bromothymol blue, and then compared one with another.

Ozonized acetone. CHEM. FAB. VON HEYDEN A-G (Georg Schroeter, inventor). Ger. 495,021, Nov. 18, 1927. A product contg active O is prepd by treating acetone with ozonized O or ozonized air until an acid reaction has developed. The product is useful as a reagent for distinguishing boiled from unboiled milk, since a drop of the product added to milk that has been treated with ext. of guaiacum wood or resin gives a blue color with raw milk but no color with boiled milk.

Apparatus for pasteurizing milk in bulk. CHARLES E. NORTH (to Creamery Package Mfg. Co.) U. S. 1,791,511, Feb. 10. Structural features.

Emulsifying apparatus suitable for use with milk and butter. L. S. RICHARDS, E. M. RICHARDS and CREMAC MARKETING CO., LTD. Brit. 335,949, April 4, 1929. Mech. features.

Apparatus and system of operation for continuous freezing of food materials such as ice cream. CLARENCE W. VOGT (to Vogt Instant Freezers, Inc.) U. S. 1,791,772-3-4, Feb. 10. Mech. features.

Butter and margarine. UNILEVER N. V. Fr. 694,770, April 30, 1930. Butter and margarine of good keeping qualities are made by bringing the H-ion concn. of the aq constituent to above 5.

Margarine. J. D. RIEDEL-E. DE HAEN A-G. Ger. 516,119, July 28, 1926. The incorporation of lecithin and other phosphatides into margarine is effected by dissolving

the phosphatides in oxidized or polymerized oils, and incorporating the soln with the margarine. Thus, lecithin 1 part and an oxidized unsatd oil 4 parts may be dissolved in ether, the soln evaporated and the product incorporated into margarine.

Cheese. SWIFT & CO. Ger 513,666, Jan 6, 1925. Emulsified cheese is produced by treating the pulverized cheese with an emulsifying agent, sterilizing and finally emulsifying without heating.

Conserving meat, etc. OTTO SANDL. Ger 513,716, Oct 27, 1927. The meat, etc., is treated with the condensed products passing over from a wood distn at about 400°, the tar oil, MeOH and acil having been removed.

Cake, pastry, etc. POMMER WERKE G M B H. Ger 516,363, Nov 2, 1929. The deterioration of cake, pastry, etc., into a dry and stale condition is retarded by thoroughly incorporating pectin into the cake, etc., mixt. Examples are given.

Jams, marmalade, etc. A W BRACH. Brit 336,321, July 22, 1929. In making jams, marmalade, preserves, confectionery, etc., the sugar is added in the form of an aq soln (preferably contg 40% or more sugar) obtained by lixivation or diffusion treatment of the sugar beet, followed by concn and purification and decolorization.

Tea extracts, etc. O STRONICH and P A WICKMANN. Brit 335,901, July 1, 1929. Tea leaves are treated with an org solvent (such as petroleum ether, pentane, CHCl_3 or ether), the solvent is removed to leave a dark green paste, and from the latter the chlorophyll, fat and wax like components are removed by treatment first with 70–80% alc, and then with a solvent such as petroleum ether, and a tea aroma residue is obtained by distg off the last mentioned solvent. A fat, oil or wax may also be used as a solvent by allowing the tea leaves to stand with it for a sufficiently long time. The tea paste may be dild with sugar to deposit the tea aroma on the latter, and caffeine and tannic acil exts may be added, prepd from tea leaves remaining in the extrn app. Tablets may be prepd from the product and it may be mixed with candy, chewing gum, milk or other edible substances.

Apparatus for drying tea. J. A. MATY. Brit 336,619, June 14, 1929. Structural features.

Alimentary salt. P WOOD. Brit. 336,279, Feb 18, 1929. Solns. contg Mg salts are added to solid NaCl to give proportions similar to those of salts from sea water and the wet mixt is dehydrated as described in Brit 319,203 (C A 24, 2513). Org Mg salts may be used and some of the Mg may be in the form of insol salts or finely divided or colloidal material.

Food for animals. R A LEGENDRE. Brit 338,014, Nov 3, 1928. Preservation of feeds such as straw, bran or oats mixed with molasses is facilitated by adjusting the pH of the contained moisture to a value between 6 and 11 by the addn of various substances effective for this purpose.

Preserving green fodder. GUSTAV FINGERLING. Ger 513,557, Nov 20, 1925. Green fodder is treated with a Cl evolving electrolyte, e.g., NaCl soln and subjected to electrolysis. The cathode is surrounded by a diaphragm. App is described.

13—GENERAL INDUSTRIAL CHEMISTRY

HARLAN S. MINER

European laboratories for the testing of materials. FRANKLIN L. EVERETT. *Mech. Eng.* 53, 201–4 (1911). E. H.

Drying the output of the chemical plant. RALPH BECKER. *Chem. Met. Eng.* 38, 91–6 (1931)—A survey of commercial drying equipment and applications.

L. W. T. CUMMINGS
The heating and drying of granular materials by convection. W. GILBERT. *Engtneer* 150, 500–2, 550–2, 574–5, 617–9, 640–2 (1930)—A study of the rate of heat transfer from hot gases to granular lumps of coal, cement, clay, etc., under varying conditions of lump size, gas velocity and gas temp. Graphs and formulas are given from which the rate of drying over a wide range of conditions can be calculated. D. B. DILL.

The economy of hot-air drying with regard to the highest permissible temperature of the product to be dried. ODOV VAJNA. *Chem. Fabrik* 1930, 25–7; cf. C. A 24, 3066.—The results of calcs. of the heat consumption and power requirements for air circulation in parallel and counter current driers and the lowering of the vapor pressure of hygroscopic materials as a function of the water content are shown on 4 charts.

J. H. MOORE

Hydrogen in modern industry. RENE DUBREIL. *Science ind* 1929, 323-4.

E. J. C.

Dust and air filtration. R. C. ROWE. *Can Mining J.* 52, 192-6(1931)—The advisability of filtering air for various machines is emphasized. Types of dusts and advantages gained by filtration are pointed out. The amount of filter surface is important. W. H. BORTON

The theory of rectification in packed columns. E. KIRSCHBAUM. *Chem. Fabrik* 1931, 38-40, 51-3, 63-5, cf. *C. A.* 24, 4194—A math method is developed for calcg the height and diam of packed towers, which takes into account the phys properties of the mixt and quantities rectified. It has the advantages that in most cases the results may be obtained graphically, and a chart shows curves for detg the height of intermittent and continuous towers. J. H. MOORE

Moisture in technical gases. II. Application of moisture calculations. FRIEDRICH LUTH. *Arch Eisenhüttenw* 4, 185-92(1930); cf. *C. A.* 24, 2814—After a discussion of the practical application of moisture calcs., various methods for measuring the moisture content of gases are described, among them the wet and dry bulb method, the pressure temp method, direct weight method and Haar hygrometer. H. STORATZ

Chemical warfare. MICHELE GIUA. *Rivista chim. sci. ind.* 1, 170-8, 193-201, 240-51, 266-72, 289-306, 321-60(1930)—Chem and phys. properties, toxicity, uses and manuf of poison gases used in the World War are described. E. M. SYMMES

Apparatus for defense against toxic gases. R. MARCILLE. *Compt rend.* 192, 382-4(1931)—The principle of the app is as follows: respiration takes place in a closed bell, the expired air is passed into a soda soln which absorbs the CO_2 , the diminished vol of air being supplemented by pure O_2 from a small tank of compressed gas. Such an app weighing 10 kg maintains comfortable respiration for one person for 2 to 3 hrs. The app is inexpensive and light in wt. and may also be made of sufficient size to protect several people at the same time. ALBERT L. RAWLINS

Chemical reactions in the oxygen isolation mask. ALBERT HLOCH. *Z ges. Schuss u. Sprengstoffe* 25, 379-81, 422-5(1930)—A discussion of the chem reactions involved and the chem problems to be solved in the design of a practical isolation-type gas mask. As fillers the alkali superoxides do not furnish sufficient O_2 , and hence a small auxiliary O_2 container is necessary. Other principal objections to this type of filler are (1) the high temp of the air entering the lungs and (2) the high breathing resistance of the canister. The alkali superoxides discussed are Na_2O_2 , K_2O_2 , and a mixt of the two, or KNaO_2 . A newer type of filler contains alkali chlorates (cf. *C. A.* 24, 5086). The reaction which furnishes O_2 is started by a type of ignition which results in an initial burst of O_2 which then continues at a rate sufficient for the needs of respiration. This filler is free from the objections cited against the alkali superoxide fillers. A. L. KIRKLE

Burning accident from sodium chlorate. J. W. DAAM. *New Zealand J. Agr* 42, 55(1931)—Clothing worn during spraying with NaClO_3 for weed eradication should be washed free of all traces of the salt before being exposed to an open fire. In one case, failure to observe this precaution resulted in severe burns on the lower limbs. K. D. JACOB

Insulation for high temperatures. E. F. ZEINER. *Power* 73, 104-6(1931)—Thermal conductivities of furnace-insulating materials and of insulating cements are compared. D. B. DILL

The thermal properties of ammoniates and analogous compounds and their use in absorption refrigerating machines (PLANK, VAHL) 2. Heat transfer in liquid NH_3 (HIRSCH) 2. Determination of the local strength of fabrics, felts, paper, etc. (SCHUBERT) 1. Solid CO_2 from Mexico (MARTIN) 18. Porous material for heat insulation, etc. (Brit. pat. 336 318) 18.

Annuaire de la Federation des Industries Chimiques de Belgique. Les produits chimiques belges. Brussels. The Federation. 112 pp.

An Index to Acts of Parliament and Statutory Rules and Orders Affecting the Chemical Industry. London. The Assoc. of British Chemical Manufs. 24 pp 2s. Reviewed in *Chem. Trade J.* 83, 170(1931).

Berl-Lunge Chemisch-technische Untersuchungsmethoden. 8th ed., revised and enlarged. Edited by ERNST BERL. Band I. Berlin. Hirschwaldsche Buchhandlung. 1260 pp. Bound, M 93.

Die chemische Industrie des Deutschen Reiches, Jahrgang 1930-1931. Band IX.

of "Deutsche Wirtschaftsbücherei" Berlin: Börsen- und Finanzliteratur, A-G Bound, M, 25 Reviewed in *Chimie & Industrie* 25, 262, *Plastics & Molded Products* 7, 153(1931).

GRAFE, V.: *Kaufmannische Grundlagen der Warenkunde und Warenkenntnis*. Band 1, Halbband 1 of "Grafe's Handbuch der organischen Warenkunde." Stuttgart: C. E. Poeschel Verlag 621 pp. M 39 Reviewed in *Chimie & Industrie* 24, 1285 (1930)

Gran enciclopedia de química industrial. Química de Muspratt. Suplemento II. Barcelona: Francisco Seix 332 pp. Ptas. 21 60, bound, Ptas. 30 50 Cf. *C. A.* 24, 2314

KIEFER, PAUL J., AND STUART, MILTON C.: *Principles of Engineering Thermodynamics* New York: John Wiley & Sons 545 pp. \$4 50 Reviewed in *Mech. Eng.* 53, 243(1931)

Separation of gas mixtures. *GES. FÜR LINDE'S EISMASCHINEN A-G* *ANT GASVERFLÜSSIGUNG* Ger. 513,683, Apr. 20, 1929. Addn. to 490,878 (*C. A.* 24, 2214) Gas mixts., such as air, are resolved into components by liquefaction and fractional evapn. in an app. such as described in 490,878, but in which the heat-exchangers and storers are modified Cf. *C. A.* 24, 5117.

Gas reactions or purifications. *COMPAGNIE INTERNATIONALE POUR LE FABRICATION DES ESSENCES ET PETROLES*. Fr. 694,997, Aug. 2, 1929. In treating gases or vapors with catalytic or purifying agents a cylindrical or annular chamber closed at both ends is used, in which the catalyst or purifying agent is placed between 2 perforated partitions which allow a free space at each end for the gas. Perforated tubes for the admission of a reactivating agent and perforated tubes for the exit of the gases formed during the reactivation run through the mass parallel to the axis of the chamber.

Separation of dust from gases. *WOODALL-DUCKHAM (1920), LTD., and A. M. DUCKHAM* Brit. 336,201, June 6, 1929. Sepn. of suspended matter from a gas, such as that resulting from the passage of air through a bed of coal and shale for stratification and removal of impurities, is facilitated by adding a gas or vapor such as steam or a liquid spray and so altering the relative ds. of the gas and dust and in some cases causing cohesion of the dust by wetting. Various details of procedure and control of the treatment are described, including an automatic control for the admission of the steam or other treating agent.

Separating fatty or waxy substances from solutions. *DEUTSCHE GASOLIN A-G* (Johann K. Pfaff, Kurt Böttger and Adolf Siewele, inventors) Ger. 516,401, May 17, 1927. See Fr. 677,385 (*C. A.* 24, 3110).

Emulsions. *DEUTSCHE HYDRAULIKWERKE A-G* Ger. 513,813, Jan. 17, 1929. Addn. to 512,979 (*C. A.* 25, 1303) In emulsifying substances insol. or only slightly sol. in water, the emulsifying agent of 512,979 is replaced by the reaction products of the esters of monobasic org. acids and alcohols of high mol. wt., with polybasic org. acids. Thus, oleic acid is esterified with AcOH in the presence of H_2BO_3 and the product treated with adipic acid. The product is used to emulsify peanut oil or other neutral oils in the presence of a small quantity of NaOH . Another example is given.

Moistening pulverulent materials. *ORTOKAR URBASCH*. Austrian 120,160, June 15, 1930. A uniform and still pulverulent product is obtained by treating the materials with the vapor or spray of the moistening liquid in a rotary drum having loose or fixed subdividing means. The treatment of stone dust, foundry sand, blue dust, fine ores and cement is referred to.

Separating lump materials according to their behavior toward rays of short wave length, including Röntgen, cathode or ultra-violet rays. *FRIED. KRUPP GRÜSONWERK A-G* Ger. 516,130, July 27, 1927.

Sterilizing closed vessels. *ARTHUR VONDRAN*. Ger. 513,632, Mar. 27, 1928. A reversible air or gas current moistens and carries around the disinfecting agent.

Canister construction for breathing apparatus. *R. H. DAVIS*. Brit. 336,674, July 19, 1929.

Use of sodium peroxide and similar compounds in cartridges of respiration apparatus. *DEUTSCHE GASGLÜHLICHT-AUER GES. and HANSEATISCHE APPARATEBAU GES. VORM. L. VON BREMEN & Co* Brit. 337,170, Nov. 7, 1928. In app. of this character, of closed circuit type, evolution of O is effected at the commencement of breathing by the introduction into the app. of a gaseous acid such as CO_2 or HCl (which may be generated from an effervescent powder or supplied in suitable flasks). Various details of app. are described.

Refrigerating agents. *K. BECK*. Brit. 336,557, June 11, 1929. In a refrigerat-

ing system employing in open circuit an aq. NH_3 soln., the NH_3 vapors are absorbed on exhaustion in a liquid such as H_2SO_4 , which converts them into a non-volatile form. Various details of app. and operation are described.

Absorbent for use with ammonia in refrigerating apparatus. HARRY F. SMITH (to Frigidaire Corp.) U. S. 1,791,515, Feb. 19. A mixt. of SrCl_2 and LiNO_3 is used as an NH_3 absorbent.

Insulators. I. G. FARRENTINO A.-G. Fr 694,151, April 18, 1930. Elec. conduct. ing wires are insulated by filaments obtained by the hardening of polymerization, condensation, or transformation products of tetradene hydrocarbons no longer elastic or extensible.

Insulators. I. G. FARRENTINO A.-G. Fr 694,425, April 25, 1930. The condensation products obtained from ethers or thioethers of dicarboxylic acids and polyhydric alcoh. described in Fr 655,755 (C. A. 24, 6041) are used as insulators in elec. machines and app.

Insulating material. GUSTAV OEDER. Fr 704,544, April 28, 1930. A sound and shock insulating material is composed of a vulcanized oil mixed with MgO , cumar., etc.

Heat-insulating material. I. G. FARRENTINO A.-G. Brit. 336,163, March 11, 1929. See Fr 691,958 (C. A. 25, 1394).

Waterproofing materials and manufacture of electrical insulation. R. C. NEVILLE. Brit. 336,731, Aug. 29, 1929. Articles such as those of asbestos, asbestos cement, clay or the like are impregnated with coal tar, tar oil, pitch or similar material, dried and immersed in a soln. of CH_3O and phenol, and finally heated to 220° for 3-12 hrs. Various details and modifications of procedure are described.

Heat- and sound-insulation. M. CASER. Brit. 336,440, Oct. 24, 1928. The residue of the combustion of rice chaff is used (suitably in the proportion of 50% or more) with other materials such as asbestos, "silicate wool" and a small proportion of binding substances such as starch or fused silicates, or with clay and granulated cork, etc. Cf. C. A. 24, 5887.

14—WATER, SEWAGE AND SANITATION

EDWARD BARTON

Potable and industrial water 1924-1929. Advances, processes, literature. H. BACH. Chem.-Ztg. 55, No. 17, Fortschrittsber. No. 1, 1-29 (1931). E. H.

Water supply conditions at Sweetwater, Texas. H. N. ROBERTS. Proc. 12th Texas Water Works Short School 27-30, U. S. Pub. Health Expt. Abstracts 11, W, 15 (Jan. 31, 1931).—The new Butter Creek supply of $2\frac{1}{2}$ million gals. per day plus a ground supply of 2 million gals. furnishes water of good quality for 11,000 people. Rapid filtration is used. C. R. FILLERS

Water works of the City of Mexico. EDUARDO MELINA, ANTONIO CORNEJO AND ALFONSO V. ACOSTA. State Inst. Water Works J. June, 1930, 11-7; U. S. Pub. Health Expt. Abstracts 11, W, 7 (Jan. 24, 1931).—The source of the water is at Xochimilco Springs, about 20 miles south of the city. Because of previous occasional contamination, a dosage of 0.1 p. p. m. of Cl_2 is now added. The present supply is to be enlarged to a capacity of 100 million gals. per day. C. R. FILLERS

Development of ground water supplies. R. O. JOSELYN. 5th Ann. Rept. Missouri Water and Sewerage Com., 64-71, U. S. Pub. Health Abstracts 11, W, 15 (Jan. 31, 1931).—Of 200 water supplies in Missouri, 67% use ground water as a source. Drilled wells are more economical and less likely to become contaminated than dug wells. Complete sealing out of water from contaminated or otherwise unfavorable strata is essential. C. R. FILLERS

Analysis of the water and deposits from the springs of the "baths of Bama" (district of Tamaulipas-Mex.). J. DICK. Bul. soc. chim. Mex. 5, 237-42 (1930).—Analytical data are given. J. H. REEDY

The nitrogenous compounds in sea water—methods for their study. J. GRUBER and coworkers. Meded. [2], 25, No. 10, 213-12 (1930).—Attention is called to the wide variety of types of nitrogenous compds. present in sea water. A review is given of the existing methods which might be applied to the study of the following classes of N compds.: free N , NH_3 , nitrates, albumins, total N and org. N compds. other than those mentioned. No expl. data are given and no new methods are suggested. JOHN B. ENDRICK

The Hardershof water-purification plant at Königsberg. G. SATTLER AND R.

BRUCHER. *Gas u. Wasserfach* 74, 7:1 6, 101-4, 128-32 (1931).—A new water purification system was constructed at the Hardershof waterworks of Königsberg in 1929 and 1930, and details of the system and its operation are given, with numerous illustrations. The yellow raw water is softened with lime water and clarified and decolorized by adding 30-60 mg. alum per l., followed by rapid filtration through a quartz sand filter after a settling period. The older slow sand filtration plant is still in use. Daily detour of color, hardness, free CO_2 , pH and permanganate requirement are made before, during and after treatment. Bacterial tests also are made. A clear and almost colorless water is obtained from the badly colored ground water. R. W. RYAN.

An interesting water-purification change. OTTO RUGEL AND ALBERT WICKIEN *Chem. Ing.* 17, 20: 8 (1934). A summary of changes made in a soda lime water-softening system and the results obtained are given, including costs. M. C. ROOPRA.

Sterilization of water by metals. F. DUBREY AND P. BERGLAND *Compt. rend.* 192, 187 7 (1931), cf. *Compt. rend.* 136, 707 (1930), Lakhovsky, 188, 1030 (1929).—Previous work has shown the efficiency of Ag, Zn and Mg in destroying the microorganisms in water. The expts. here reported show that water is freed from *B. coli*, pathogenic organisms, and liquefying bacteria by passing it over a prepri called silver sand or by agitating it with the silver sand. The organisms disappear more rapidly from water at 47° than at a lower temp. The presence of NaCl in the water retards the process of bacterial destruction. It is shown that water thus treated contains less than 0.001 mg. of Ag per l. N. M. NAYLOR.

Experience with the lime-baryta process. A. LARSEN *Ing. De. Norge* 53, 913-9 (1930).—The results of a no. of investigations on water purification installations are presented. The effect of the addition of lime and baryta, and the limit of softening are dealt with. ALLEN S. SMITH.

Boiler-water conditioning. A Pittsburgh development. J. N. WELSH AND H. A. JACKSON, *Proc. Eng. Soc. West Penn.* 46, 327-42 (1930), cf. *E. I.* 23, 3307.—Success in boiler operation at high pressures depends upon the timely and correct analysis of the constantly changing boiler H_2O . Fundamental requirements of boiler H_2O conditioning establishes that a sufficient amt. of dissolved material (carbonate or phosphate) must be maintained to give chem. equilibria, all org. or saponifiable material must be eliminated, and a min. concn. of solids in feed H_2O be maintained. Numerous specific examples, from the Pittsburgh area, of such conditioning are cited. EDWARD S. HOPKINS.

Determining the quality of boiler-feed water by the electrical method. Yr. M. KOSVATSKY, *Izvestiya Vsesoyuzn. Inst. (Trans. Techn. Sch. Inst., Russia)* 1930, No. 2, 27-32.—The performance of the Leeds Northrup instrument of Philadelphia is described. A. A. BOHRNTRINK.

Purification of sea water for boilers. I. SUIK AND G. TOROGYAN, *Armenian. Khim. i Tekhn. Zh.* 1930, No. 10, 101-8.—Prevention of CaSO_4 formation is of primary importance. Ninety-five % of the Ca is pptd. by treating 100 cc. of water with 0.317 g. Na_2CO_3 . Fair results are obtained even with 0.075% Na_2CO_3 . V. KATCHEVSKY.

The study of the fouling of steam turbine paddles. J. R. BERNOLD, *Chaleng. et Ind.* 11, 178-80 (1931).—Two 4500-kw. turbines running on steam at 14-15 kg. at 320° suddenly gave trouble after some years of satisfactory performance, by blade fouling in 2 weeks so that the rotor was quite out of balance. Two analyses of the deposits are given, the major constituents being CO_2 , CaO , Fe_2O_3 and SiO_2 . It was concluded that the trouble was due to the boiler feed waters. Finally this was held at 40-45° of alk., i. e., about 2 g. per l. Some discussion is given as to how the deposits arose and also a description of the kind of corrosion seen in the steel tubes of water reheaters. S. L. B. BRIDGTON.

Practical operation of sewage works. C. R. KERRA, *et al.* *Civil Eng.* 1, 91-4 (1930).—This is a committee report and subsequent discussion before the Sanitary Eng. Div., Amer. Soc. Civil Eng. The report is a general review of sewage-plant operation and outlines principles for proper procedure. The detailed tests for control purpose are given together with suggestions as to the reporting of cost data, etc. The discussion, showing wide variance in personal opinion, was based upon the value of the various claims, and their relationship to the problems in question. EDWARD S. HOPKINS.

The simple facts about sewage treatment. G. O. G. NASHMITH *Can. Pub. Health J.* 22, 70 6 (1931).—A discussion of sewage treatment, with particular reference to the North Toronto activated-sludge plant, which treats a normal dry weather flow of about 4 million gallons per day from a residential district of 60,000 people. The plant consists of a storm water tank, bar screens, primary settling tanks, aeration tanks in which rising air bubbles give the sewage a spiral motion, final settling tanks, sludge-digestion

tanks and glass covered drying beds. The effluent is non putrescible and contains 15-20 p p m suspended matter, its quality being better than that of the Don River, into which it flows. The primary and activated sludges are continuously digested together, the 30,000 cu ft. of gas collected being employed for heating the incoming sludge and for lab purposes. Digestion is complete in a month. The digested sludge contains 50% org matter and approx 3% N₂ and is odorless. R. E. THOMPSON

Sewage-reclamation plant for Los Angeles. R. P. GORDY. *West Construction News* 5, 110 25(1930). *U S Pub Health Eng Abstracts* 11, 8, 10(Jan 17, 1931).—In order to study sewage reclamation, the City of Los Angeles placed in operation in May, 1930, an exptl sewage plant of 200,000 gal-per day capacity. The problems to be studied are outlined. C. R. FELLERS

Modified form of Imhoff tank used in four Oklahoma towns. FRANK B KING. *Eng News-Record* 106, 112(1931).—General practice of allowing 15-20% of top area for gas vents is probably satisfactory where there is daily supervision or mech means is employed to break up the scum, but experience has indicated that a greater area must be allowed for gas vents in installations where little attention is paid to operation. In modified type of tank employed in 4 towns in Oklahoma, 60% of the top is left for scum accumulation, the sludge bed being placed nearby to allow the scum to be thrown from the tank with a fork when excessive amts interfere with escape of the gas. R. E. T.

Thermophilic digestion of sewage solids. IV. Fresh solids and activated sludge. WILLIAM RUDOLPH AND H. H. FUKLEKIAN. *Ind Eng Chem* 23, 67-70(1931), cf C A 25, 157.—Digestion of lab samples of daily addns of fresh and activated sewage solids was accomplished at 60° in the short period of 2 1/2 days, seeding materials produced under similar thermophilic conditions having been used. These expts. indicate that a daily charge of 50% volatile matter as fresh sludge can be assimilated. The sludge so obtained had all of the characteristics of well-digested material. Additional work is to be conducted on a semi plant scale. EDWARD S. HOPKINS

Dewatering sewage sludge on glass-covered sand beds. G. C. HOUSSER. *Muniz Sanit* 2, 69-70(1931).—Glass-covered beds afford means of at least partially controlling odors from drying sludge. By protecting the sludge from rain the area can be reduced as much as 50%. D. L. KELSO

Separate sludge-digestion plant at Toledo mechanically operated throughout. HARVEY P. JONES. *Eng News-Record* 106, 389-92(1931).—A description is given of what is stated to be the world's largest mechanically operated sep sludge-digestion plant, now under construction at Toledo, Ohio. Treatment will include prechlorination, coarse screening, grit removal, sedimentation, sep sludge digestion and sludge drying. Mech equipment will provide for automatic cleaning of coarse screens, continuous cleaning of grit chambers, continuous sludge removal from sedimentation tanks, skimming of scum and oil from the latter, and agitation of sludge during digestion. The plant, which is the final step in a 14 year program for cleaning up the Maumee River, is designed to serve the 1933 population of 600,000, contributing a mean daily flow of 80 million gallons. Prechlorination will be employed to eliminate odors and reduce the bacterial load. The chlorination plant being designed for an ultimate capacity of 18,000 lb of Cl₂ per day, equiv to a dosage of 15 p p m based on the max dry weather flow of 140 million gallons per day anticipated in 1950. The screenings will be destroyed at the municipal incineration plant. Digestion tank capacity equiv to 223 cu ft. per capita will be provided, including 6 months' sludge storage during the winter. The gas formed will be collected in domes and utilized for maintaining the optimum digestion temp. Provision is also being made for maintaining the pu value of the sludge at the optimum for digestion. The drying beds will be glass-covered, the space provided being equiv to 0.25 sq ft per capita. The dried sludge will be used for reclaiming low lying ground. The cost of the plant was estd at \$1,971,000 and the actual bids were approx 5% lower. R. E. THOMPSON

The influence of chlorine on an activated-sludge process. H. D. BELL. *The Surveyor* 79, 217-8(1931). cf C A 24, 1450.—Since bleaching powder is cheaper than liquid Cl for small plants, it was used to reduce the load on the bio-aeration unit. Twenty % increase in plant capacity resulted from chlorination. The final effluent showed an av of 2.0 p p m of nitrate N, thus showing that putrefaction was decreased by Cl treatment. A. L. FLORE

Filtering materials for trickling filter. WM. E. STANLEY. *Civil Eng* 1, 113-6(1930); cf C A 25, 368.—A general review of the materials used in sewage filter beds. EDWARD S. HOPKINS

Selecting the materials to prevent clogging. FRANK W. JONES. *Civil Eng* 1, 118-20(1930).—The size and durability of filter-bed material must be such as to insure

an abundant air supply through the entire mass. Surface growths seal the beds, diminish the air supply and cause "ponding." This is also caused by application of improperly treated sewage. Disintegration of the beds may cause clogging in the depths, decreasing the voids. This may also be caused by the rough surfaces of the aggregates retaining the solids. Uniform particles of clean material should be used, regardless of size. They must be as durable as possible to retard disintegration by weathering.

Operating filters to remove organic growths. CHAS. C. HOMMON, *Civil Eng. I.*, 116-8(1930).—Study of about 60 sewage trickling-filter beds in this country shows that while reasonably clean filters are obtained naturally it is doubtful if the process is satisfactory. The uncertain natural removal of clogged material by sloughing, entirely beyond the control of plant operators, is a disquieting feature of plant control. Control of sloughing and killing of surface algae was accomplished by drying the beds for 24 hrs. every 7 days. This eliminated "ponding" and cleaned the top 12 in. of material of adhering particles by action of the sewage flow when re-admitted to them. The results of this practice in the Canton, O., plant were in accordance with expectation and prevented the more or less periodic "ponding" of the beds.

Experiments on the destruction of the filter fly (*Psychoda*). W. D. SCOLLER AND H. H. GOLDBROOK, *The Surrey* 70, 219-20(1931).—Expts. were made on 2 filters with creosote oil emulsion and *o*-dichlorobenzene. Expts. in 1. cylinders showed that vapors from *o*-dichlorobenzene killed the flies in 10 min. Treatment on filters at the rate of 1 gallon per day per 437 sq. yds. of filter surface inhibited breeding of *Psychoda* after 10 days' treatment and gave immunity for a month. It was twice as effective as the creosote emulsion in paraffin oil. One gallon of creosote emulsion per day was needed per 220 sq. yds. of filter surface. Flies gradually returned when the treatment was stopped.

Notes on the determination of biochemical oxygen demand. W. D. HATFIELD AND K. MORKERT, *Sewage Works J.* 2, 521-8(1930).—The diln. method was found to be accurate to 5% if the bottles were carefully cleaned with cleaning soln., rinsed and dried, and if the org. matter in the diln. water was well stabilized and the water itself antitoxic to the normal bacterial flora. Distd. water for diln. gave low results, 500 p. p. m., bicarbonate water higher results, and natural well water still higher results. Adoption is recommended of a standard diln. water congl. Ca, Mg, Na, HCO_3 , SO_4 and Cl ions. Data are given which show that rubber stoppers have an O demand and should never be substituted for glass stoppers in the incubation bottles.

Effect of sunlight and green organisms on regeneration of streams. WILLIAM RUDOLPH AND H. HEUKELKIAN, *Ind Eng Chem* 23, 75-8(1931), cf. *C. A.* 23, 4760.—This lab. study is based upon the dissolved O, biochem. O demand and B. coli tests of samples collected from the Delaware River for 24 hr. periods. Results, presented in chart form, indicate that the dissolved O increased rapidly during the forenoon, reaching a max. between noon and 4 p. m. During the night a decrease was noted until the lowest point was reached at 3 A. M., with a subsequent increase again after daylight. This occurred in spite of temp. decrease during the night hrs. Many expts. confirm the belief that this O is given off by blue green algae growths in vigorous activity. These data have important bearing on stream-pollution surveys, particularly in summer seasons and discredit afternoon sampling if these specimens of algae are present.

Technic of stream-pollution investigations. F. W. MOULMAN, T. L. HERRICK AND H. GLADIS SWORN, *Ind Eng Chem* 23, 209-13(1931).—Special considerations are to be noted in the sampling of a polluted river. A carefully chosen sampling program is to be followed and the most representative points in the stream used for the testing stations. Bottom samples should be carefully collected to assure the elimination of sludge deposits. Composite samples are a necessity for correct data. Interesting diagrams locating proper sampling stations are given.

Experimental studies of natural purification in polluted waters. IV. Influence of the plankton in the biochemical oxidation of organic matter. C. T. RUTHERFIELD, W. C. PURDY AND R. J. THURIAULT, *U. S. Pub Health Repts* 46, 303-426(1931); cf. *C. A.* 24, 2220.—A dil. dextrose-peptone soln. (D), which could be readily and accurately reproduced, was used in a series of expts. designed to show the functions of the bacteria and the plankton and the probable interrelationships of the 2 groups of organisms in the biochem. oxidation process. The results obtained in these expts. indicated that: (1) Dil. I preserved free from biol. activity did not absorb any dissolved O under the conditions of these tests. (2) When inoculated with bacteria in pure culture, dil. I favored their growth, and they increased rapidly in no., reaching a limiting population by the second day of incubation. This limiting no. was maintained for long periods of

time (3) While bacteria were actively multiplying, O was depleted at a rapid rate. After the limiting no. had been reached, this depletion of dissolved O practically ceased, although the living bacterial population remained high (4) The results obtained with mixed cultures of bacteria, free from plankton, were the same as with pure cultures, except that the extent of oxidation was somewhat greater (5) The protozoan *Colpidium* grew well in the presence of bacteria in dil I but was not able to grow in it in the absence of bacteria. When the concn of food in the medium was increased 100-1000 fold *Colpidium* grew well in the absence of bacteria. The conclusion was reached that in the dil medium the bacteria act as "collectors" or "concentrators" of the *Colpidium* food (6) *Colpidium* growing in pure culture used up the O. The amt. of O used, however, was comparatively small (7) When bacteria and plankton were grown together in dil I, the results obtained during the first 2 days of incubation were approx the same as when bacteria only were present. After the first 2 days, however, the bacterial nos. were not maintained but were reduced rapidly, the reduction being accompanied by a plankton increase. Moreover, the oxidation process did not cease but continued as in natural polluted waters (8) In general the extent of oxidation observed in dil I varied directly with the complexity of the biol. factors present—that is, the greater the variety of organisms acting in the medium, the more extensive the O depletion observed. Based on the results obtained the theory is advanced that the chief function of certain plankton in the biochem. oxidation process is to keep the bacterial population reduced below the satn. point and thus to provide conditions suitable for continuous bacterial multiplication. This in turn resulting in more complete oxidation. Support is given to this theory by the results obtained in expts. in which the limiting nos. of bacteria were reduced by phys. and chem. means. Such reductions in bacterial nos. were invariably followed by renewed bacterial multiplication and oxidation. J. A. KENNEDY

Some inter-relationships of plankton and bacteria in natural purification of polluted water. C. T. BUTTERFIELD AND W. C. PURDY *Ind. Eng. Chem.* 23, 213-8 (1931), cf. C. A. 22, 2221.—Normal polluted H₂O contains many bacteria and plankton, if these are removed or killed oxidation ceases. Studies with *B. aerogenes* and bacteria free *Colpidium* in pure culture were undertaken to show the relationship between them. Active multiplication of bacteria caused rapid adsorption of O by the H₂O until a limiting concn was reached. Use of the pure plankton culture gave moderate O depletion, admixt. of the cultures gave rapid O adsorption during periods of bacterial multiplication and also after the limiting value of *Colpidium* had been reached. This indicates that the function of plankton in polluted streams is to maintain the bacterial population below their limiting value and to provide suitable conditions for their multiplication, thereby giving more complete oxidation. EDWARD S. HOPKINS

A routine method for direct determination of *B. coli* in large quantities of water on a solid medium. I. DIENERT AND P. ETRILLARD *Ann. sciences Tech. hyg. Ville de Paris* 1930.—In France the common test for *B. coli* is by the use of 0.1% phenol broth. The medium described is a type of eosin methylene blue agar. It contains 1000 ml. distilled water, 10 g. peptone (Difco), 2 g. dipotassium phosphate, 10 g. lactose, 15 g. agar. The reaction is adjusted to pH 7.5. When ready to use the following is added to 25 ml. of the agar: eosin yellow 2%, 0.5 ml.; methylene blue 0.5%, 0.5 ml.; phenol 5%, 0.25 ml. Highly contaminated waters are plated directly in small quantities, in waters less densely polluted the bacteria must be concd. *B. coli* colonies are round, and 3-4 mm. in diam. By transmitted light the center of the colony is a dark violet blue, this color covering at least 1/4 of their area. The periphery is bright gray blue, sometimes metallic in appearance. By reflected light the center of the colony appears slightly elevated. The colonies beneath the surface are in the form of small, dark blue lentils. *B. aerogenes* colonies are 4-6 mm. in diam. with deep-brown centers when viewed by transmitted light. By reflected light they do not show the metallic sheen. In case of doubt colonies are picked and inoculated to peptone broth. G. L. KELSO

Epidemic of intestinal disorders in Charleston, W. Va., occurring simultaneously with unprecedented water-supply conditions. E. S. TISDALE *Am. J. Pub. Health* 21, 198-200 (1931).—Throughout the entire drought period of 4 mos. (Aug. to Dec., 1930), in which acute and offensive taste and odor conditions existed in the water supply and a high pollution continued as a result of pollution of the Elk River by water from the Kanawha River, the filtration plant operation was above reproach, and daily independent bacteriol. tests by 3 different water testing labs. showed that the U. S. Treasury Standards had been met. The illness caused by these unprecedented circumstances is described. J. A. KENNEDY

Water-borne typhoid fever still a menace. ABEL WOLMAN AND ARTHUR E. GORMAN *Am. J. Pub. Health* 21, 115-23 (1931).—W. and G. give a detailed analysis of water-

born* outbreaks of typhoid fever and *dysentery* covering the decade 1923-1929 for both the United States (47 states) and Canada (6 provinces). These outbreaks are classified as to population groups affected, as to points of pollution in waterworks system, as to causes and as to relative magnitude. Data are given concerning the largest outbreaks. The 5 cities in the U. S. with repeat outbreaks from the same cause and the 4 cities with repeat outbreaks from a different cause are further considered in tabular form.

J. A. KENNEDY

Use of chloramine in treatment of pool water. J. F. T. BERLINER. *Beach and Pool Mag.* 5, 9-10 (1931).—Chloramine unlike chlorine, is not destroyed by org. matter or actinic rays, nor removed by aeration. Less chlorine is required, a reduction as high as 80% being accomplished. Complaints of tastes, odor, and irritating effects, as well as algae and slime growths are reduced. The ammonia is added before the chlorine. Before applying the o-tolidine test, nitrites, if present, must be destroyed by oxidation with H_2O_2 .

G. L. KELSO

Chloramine, its preparation, properties and use. HIR-CHEIND' 6. Importance of microbiology [in study of sewage and water]. CAZI 15. Apparatus for drying sludge for fertilizer (U. S. pat. 1,792,064). Recovering fbers from the waste waters of the paper, cellulose, etc., industries. Austrian pat. 1,296,110. 23. Apparatus for eliminating CO from motor vehicle exhaust gases (U. S. pat. 1,791,921). Purifying light oils recovered from coal gas, etc. [with production of solutions for softening water] (Brit. pat. 336,636) 22.

Regenerating zeolites used in water softening. WALTER H. GREEN (to Permutit Co.). U. S. 1,792,220, Feb. 10. A moving stream of the zeolite is first subjected to the action of a relatively weak regenerating soln. and then to the action of a stronger regenerating soln. App. is described.

Tubular heat-exchange apparatus suitable for heating boiler feed water, etc. FOSTER WHEELER CORP. Brit. 336,859, Jan. 23, 1930. Structural features.

Dechlorinating water. JOSEF MUCIKA. Austrian 629,119, June 15, 1930. In removing excess of Cl from chlorinated water by passing the water through active C or the like, the dechlorinating mass is kept sterile by continuously exposing a fresh surface to the chlorinated water. Suitable, the water is supplied through an injector to the top of a container for the active C, and the water supply pipe is connected to the bottom of the container in such a way that C is sucked up by the injector and carried by the water to the top of the C mass.

Rotary sewage distributor. E. HARTLEY. Brit. 336,384, Sept. 23, 1929. Structural features.

Purifying waste waters from gas manufacture. HEINRICH BRENS. Ger. 516,527, July 23, 1929. Addn. to 504,337 (C. A. 24, 5404). The washing oils required for the method of Ger. 504,337 are used in a cycle and are dephenolized, e. g., by washing with NaOH soln., immediately before or after treating the waters therewith.

15—SOILS, FERTILIZERS AND AGRICULTURAL POISONS

J. J. SKINNER AND M. S. ANDERSON

The formation of a board of specialists to study the interesting scientific problems of agriculture in the hot countries and the organization of scientific service for the Colonies. AUG. CHEVALIER. *Bull. agr. Congo Bdg.* 21, 355-73 (1930).—The formation of a board to study the agricultural and forestry possibilities of the hot countries and to furnish scientific information is discussed. A no. of the important problems facing such a comm. are given.

J. R. ADAMS

Importance of microbiology as a subject of study and research in agricultural colleges. A. K. B. CAZI. *Portug. Agr. Col. Mag.* 22, 172-9 (1930).—The importance of microbiology in the study of soils, manure, sewage, water and, particularly, milk and its products is discussed.

K. D. JACOB

A few impressions of agricultural chemistry in foreign countries. C. A. BROWNE. *J. Assoc. Official Agr. Chem.* 14, 101-12 (1931).—An address dealing briefly with modern trends in European countries, as observed in the course of 16 months' sojourn in various European countries.

A. PAPINEAU-COUTURE

The restoration of the fertility of tropical soils. ALBERT BRUNO. *Bull. agr. Congo Bdg.* 21, 574-80 (1930).—A no. of the plantations in the French colonies have

lost much of their fertility because of lack of fertilization. Conclusions are drawn from work carried out in three countries as to the proper fertilizer for the various tropical crops on these plantations. J. R. ADAMS

Titanium in some New Zealand soils and pastures. H. O. ASKEW. *New Zealand J. Sci. Tech.* 12, 173-9(1930).—Soils of various types from different regions were examined. The Ti content varied from 0.25 to 1.5%, called as Ti, but in a large proportion of the samples it fell between 0.3 and 0.7%. "Bush sick" soils, both of granitic and volcanic origin, were low in Ti. Heavy loam and clay were highest. The variation in Ti content of the pasture grasses and clovers is considerable. From 1 to 8.5 mg TiO₂ per 100 g dry matter was found. This amt. is believed to be too small to affect detrimentally the assimilation of phosphates in the animal body by pptn. of an insol. phosphate. There appears to be a correlation between the Fe and Ti contents of the pasture vegetation, both tend to increase at the same time. No correlation exists between Ti and Mn. L. E. GILSON

Chemical composition of the soils of McHenry County. T. H. HOPPER AND H. L. WALSTER. *N. Dak. Agr. Expt. Sta., Bull.* 240, 3-46(1930).—The soils are carefully classified as to type, depth and chem. compn. The following composite % compn. of surface soils is representative: N 0.247, P 0.000, K 1.35, Ca 1.66, Mg 0.55, S 0.91, Si 35.22, org. C 2.72, carbonate C 0.24. With the exception of the gravelly and sandy soils, sufficient Ca, S, N, P and K were present for normal crop needs. C. R. F.

Chemical analyses of Congo soils. F. CLAUS. *Bull. agr. Congo Belge* 21, 560-1(1930).—A short note is given on the history of the analyses of Congo soils with the titles of a few of the publications on this subject. J. R. ADAMS

Chemical analyses of Congo soils at the Zambé Laboratory. G. BATZ. *Bull. agr. Congo Belge* 21, 562-6(1931).—Both the mech. and chem. analyses of 30 representative soil samples of the colonies are tabulated. The soils of the central basin are very mediocre in nutritive values while in the rest of the colony these values are variable and depend on the origin of the soils. J. R. ADAMS

Response of Illinois soils to systems of soil treatment. F. C. BAUER. *Ill. Agr. Expt. Sta., Bull.* 362, 437-614(1930).—This bulletin brings together in summarized form, certain results obtained from 28 exptl. fields that are located on representative types of Ill. soils and have been under investigation for 15-26 years. The immature, dark soils of Ill. are much more productive than the mature lighter soils. Farm manure with or without lime was an effective fertilizer on all soil types. The proper use of lime is fundamental to the management of most Ill. farm lands, and nearly always gives very profitable returns. In general, soils deficient in available P respond generously to applications of superphosphate or very finely ground raw rock phosphate. Little difference was found as to the effectiveness of these 2 carriers of P. The light-colored soils and the less-productive dark-colored soils respond strikingly to applications of K. C. R. FELLERS

Microbiological studies of some typical Iowa soil profiles. P. E. BROWN AND T. H. BENTON. *Iowa Agr. Expt. Sta., Research Bull.*, No. 132, 363-420(1930).—The nos. of bacteria, actinomycetes and molds in various horizons of typical profiles of some Iowa soils were detd. during 2 seasons, and the results are compared with careful descriptions of the soil profiles by horizons. C. R. FELLERS

Studies in soil colloids. I. Base exchange and soil acidity. AMAR N. PURI. *Mem. Dept. Agr. India, Chem. Ser.* 11, 1-38(1930), cf. C. A. 25, 159.—The exptl. evidence obtained in this investigation lends support to the view that soil acidity is due to insol. colloidal acids or "acidoids," which are assoc. with surface active H ions and characteristic reactions. A no. of reactions characteristic of soil acids were studied in fully unsatd. soils obtained by exhaustive treatment with 0.05 N HCl. The results showed that acidoids, though insol., give reactions similar to those of soil acids, the reactions taking place, however, in 3 distinct stages in which 1, 2 or 3 equivs. of H can take part. The exptl. facts can be best interpreted by supposing that the acidoid in a fully unsatd. soil is a tribasic acid. Results of huming expts. showed that toxicity of acid soils is confined to the primary H ion of the acidoid. II. Factors influencing the dispersion of soil colloids in water. *Ibid.* 39-51.—Dispersion coeff. (D.C.) is defined as the % of the total clay (0.002 mm) that can pass into suspension by "auto-subdivision" on being left in contact with water for 24 hrs. Expts. with air-dried clayic acid (obtained by prolonged treatment of clay with 0.05 N HCl) and metallic clayates (obtained by exhaustive treatment of clay with neutral salts) showed that Na clayates go almost completely into suspension after air-drying by simply keeping them in contact with water, while K, Ca and Ba clayates have a D.C. of approx. 10 or less. The D.C. for clayic acid and Al clayate is very low. Na and Li clayates contain the largest pro-

portion of particles that show no sign of settling to a depth below 5 cm. after 19 days 11 a soil suspension at max. dispersion is converted, however, into any clayate (with the possible exception of the clayates of Al and Fe) the dispersion remains at its max., provided the soil is not allowed to become dry. Alkali soils rich in exchangeable Na show a high D C. and a very low rate of percolation. Successive treatments of alkali soils with gypsum increase the permeability and lowers the D C. Highly acid soils give a very low D C. and a very high rate of percolation. Irrigated soils give a comparatively higher D C. than unirrigated ones. III. Flocculation of soil colloids. *Ibid* 101-18.—Almost equiv. quantities of Ca, Ba and Sr hydroxides are required to flocculate a soil. The higher the valency of the exchangeable ion, the smaller the ionization and stability of the suspension and consequently the smaller the flocculation value (ΓV) for a given electrolyte. NaCl is an efficient flocculating agent when the exchangeable ions are principally Fe and Al, but is quite inefficient in the case of soils contg. principally exchangeable Na. In general, Na ions are the weakest flocculating agents, and they show a great difference between ΓV for partial and complete flocculation. Where the amt. of exchangeable Na is large, it is not possible to locate accurately the latter value because the whole mass gelatinizes. Flocculation values for Al ions are not greatly different from those for Ca and Ba ions. This is probably due to the formation of Na aluminates at higher pH values which behave like other Na salts as regards their flocculating power, or it may be due to pptn. of a part of the Al as $Al(OH)_3$. Mg ions are less powerful flocculating agents than Ba and Ca ions. The concn. of gypsum needed for flocculating a deflocculated soil contg. only bivalent exchangeable ions is 1.574 mg. of Ca per 100 cc. This shows that lime cannot function as a permanent cure for the deflocculated condition of a soil because the lime must be converted eventually into $CaCO_3$, a solid soln. of which contains only 0.72 mg. of Ca per 100 cc. IV. Methods of estimating soil colloids. *Ibid* 119-31.—Methods of detg. soil colloids are briefly reviewed. Heat of wetting, dye adsorption and NH_3 absorption methods are affected by the nature of the exchangeable base in the soil and are not reliable measures of the colloid content. Mech. analysis and water-absorption methods seem to be the most accurate. It is suggested that the hygroscopicity of a soil should be measured by bringing it into equil. with an atm. of 70% humidity, and soil in equil. with an atm. of 10% humidity should be considered as dry basis for referring to hygroscopic moisture. In order to define completely the colloidal behavior of a soil it is necessary to know its clay content, hygroscopicity and base-exchange capacity. V. Methods of determining saturation capacity and degree of saturation of soils. *Ibid* 133-41.—The following method is proposed for detg. the state of satn. of soils with respect to bases: To 10 g. of the soil add 100 cc. of water and shake for 2 hrs. with 2 g. of $CaSO_4$. Next add 100 cc. of 0.2 N $Ba(OH)_2$ and shake at intervals for 6 hrs. during 2 days. After addn. of 10 drops of phenolphthalein soln. the color of the supernatant liquid is discharged with 0.1 N H_2SO_4 , and the acid is then added gradually with vigorous shaking after each addn. The titration is complete when the supernatant liquid remains absolutely colorless for 2 hrs. The titration is completed in 2 to 3 days. The total $Ba(OH)_2$ minus the total acid gives the amt. of base retained by the soil and is equal to $T - S$, where T is the amt. of base the soil is capable of binding and S is the amt. of absorptively bound base originally present in the soil. T is detd. by washing another 10 g. of soil with 0.05 N HCl till all the replaceable bases are removed, as detd. by the absence of Ca ions in the filtrate, followed by washing with water till the free acid is removed. The soil is then transferred to a bottle and titrated with H_2SO_4 after addn. of an excess of $Ba(OH)_2$, exactly as described above. The method is simple and rapid, inasmuch as detns. can be made on 20 soils at one time.

K. D. JACOB

The adsorption of the anions of acid dyes by soil colloids. J. G. SMITH AND P. L. GILE. *J. Agr. Research* 41, 401-13(1930).—Several soils contg. colloids with widely different silica sesquioxide ratios were shaken with either H_2O or weakly acid solns. of the acid dyes, Na carminate, orange II and ponceau with negligible adsorption of the anions. The soils removed large quantities of ponceau from a 0.05 N acid soln. but this is evidently a pptn. of the dye.

J. R. ADAMS

A simplified method of determining "sticky point" of soils. AMAR NATH PURI. *Agr. J. India* 25, 206-9(1930).—The method of Keen and Coutts (*C. A.* 23, 5261) for detg. the "sticky point" of soils is simplified as follows: 50 g. of air-dry soil is spread on a stout glass plate and water is added gradually from a buret. After each small addn. of water the soil is well mixed and worked into a paste, first with an Fe spatula and then with the hand, exactly as in the standard method. The sticky point generally can be located within 2-3 drops of water, corresponding to about 0.4 to 0.6% moisture content. The amt. of water in cc. multiplied by 2 gives the sticky point. On 48 samples of soil

from different parts of India, the results obtained by the simplified method were almost identical with those obtained by the method of Keen and Courts if allowance was made for the air-dry moisture content in very hygroscopic soils. K. D. JACOB

Soil humus. S. OSUGI AND Y. SANO. *J. Sci. Soil and Manure; Proc. Intern. Soc. Soil Sci.* 5, 112-3(1930).—Soil humus was fractionated by Waksman's method. The soil portion is too small to be significant. The C, N and ash contents of the α and β fractions were examd. No particular relationships between the α fractions of rich humus soils or the β fractions in mineral soils were observed. Schollenberger's chromic acid method for detg. humus (C. A. 22, 291) is satisfactory if samples of suitable size are used e. g., 0.01 g. sample of soil with 30% humus, 0.02 g. with 15.20% and 0.1 g. with 5%. B. C. A.

Potassium thiocyanate method for determining soil acidity. Y. KAMOSHITA. *J. Imp. Agr. Expt. Sta., Tokyo, Proc. Intern. Soc. Soil Sci.* 5, 103-4(1930).—The color intensity in the Comber test for soil acidity was standardized by titrating the ale. ext. with KOH soln. Values so obtained for a no. of soils were compared with pH values exchange acidity and exchangeable base content. Soils exhibiting exchange acidity have $pH < 5.6$. Soils with $pH < 6.83$ give, with Comber's test, red colorations which increase in intensity with increasing exchange acidity. In soils showing exchange acidity the proportions of exchangeable Ca and Mg are smaller than those of Fe and Al. The reverse is the case where no exchange acidity exists. It is suggested that the red coloration in Comber's test is due to the exchange of ferric ions for K ions in the thiocyanate and that exchange acidity is partly the result of the presence of exchangeable Fe. B. C. A.

The effect of lime on the texture of soils. WALTER REIMER. *Miss. Arch. Londou, Abt. A, Pflanze* 4, 251-70(1930).—The heat of wetting of loam, clay or humus soils remains unchanged by addn. of CaO or CaCO₃. These substances increase the bacterial decompn. of org. matter in the soil, thereby changing the texture. This effect is negligible, however, for the quantities of CaO or CaCO₃ used in practice. W. G. R.

Lime additions to friable clay. W. BREYER. *Bull. Agricol. Inst. Finland* No. 29, 13 pp (1929), *Proc. Internat. Soc. Soil Sci.* 5, 129-30(1930).—Addn. to soil of lime in quantities predicted from titration curves to produce pH changes up to 5.6 effected the calcd. changes. To bring about pH 6 in the field, 1.6 times the quantity of lime calcd. from lab. tests was required. It was impossible in the field to bring the pH value to 7.0. The most heavily limed areas of this soil rapidly developed further acidity. B. C. A.

The effect of plowing under cane trash upon the available nitrogen of the soil. Wm. L. OWEN AND W. P. DUNSON. *Zentr. Bakt. Parasitenk.* 2 Abt., 82, 174-99(1930).—The addn. of fresh cane trash to soils contg. nitrates causes a very rapid transformation of the N into org. forms. This depressing effect on soil nitrates decreases at a fairly const. rate when the crop residue is in contact with the soil, and under ordinary conditions it has perhaps lost most of its pernicious effects at the time the fertilizer is applied to cane fields. The addn. of fresh trash retards appreciably the rate of growth of corn plants but when it has been partially buried for several months, it increases the initial rate of growth and the yield. The addn. of cane trash to soil increases its water holding capacity and retards drying. Cane trash in soil causes appreciable gains in total N, indicating stimulation of the N fixing bacteria by the products arising from decompn. of the org. matter. Field tests over a period of several years show no decrease in crop yields, indicating that the practice is consistent with conserving available N in the soil. Whether the continued practice of incorporating such large quantities of org. matter with a C:N ratio of approx. 1:40 would tend so to increase the accumulation of the more resistant constituent, lignin, as to affect unfavorably the compn. and productive ness of a soil needs further study. JOHN T. MYERS

Cellulose-decomposing soil-fungus of an unusual character. H. L. JENSEN. *Proc. Linnean Soc., N. S. Wales* 55, 699-707(1930).—A fungus, probably belonging to the genus *Botryosporium*, was isolated from an English field soil with addn. of manure. This organism proved very sensitive to acid reaction, pH 4.5 being very near the limit of acidity at which growth could be induced, a good growth would only take place at pH values above 6.0, and an optimum zone seemed to stretch from pH 6.8 to pH 7.4 and possibly higher. In neutral or alk. soln. the fungus exerted a very strong cellulose-decompg. activity, in unbuffered physiologically acid soln. it exerted almost no activity. Its sensitiveness to acidity is greater than that of any fungus hitherto studied.

Agreement between the Azotobacter and seedling plant methods (for determining the phosphoric acid requirements of soils). H. NIKLAS, H. POSCHENVAIEDER AND F. ALICE W. SPYERSON

CZIBULKA *Das Superphosphat* 6, 111-5 (1930). cf *C. A.* 24, 5410 —In 2000 comparative expts. with the *Asotoxeter* and the Neubauer seedling plant methods for detg. the available P_2O_5 content of soils, the agreement between the 2 methods was very good, good and poor in 74.7, 14.1 and 11.2% of the trials, resp. **K. D. JACOB**

The evaluation of the results of soil analyses, with respect to the need of phosphate fertilization. **ANTONIN NEMEC**. *Bull. Arch. Landw. Abt. A, Pflanze* 4, 228-46 (1930). cf *C. A.* 24, 5917 —Addn. of superphosphate to soils contg. but little phosphate sol in 1% citric acid soln failed to increase the yield of sugar beets in some cases, whereas other cases were observed where increased yields were obtained from soils rich in phosphate sol in 1% citric acid soln, when similarly fertilized. This anomalous behavior was found to be due to the Fe content of the soil. Soils contg. more than 45-50 mg sol Fe per 100 g soil failed to show increased yields when fertilized with superphosphate, regardless of the phosphate content of the soil. Other expts. showed that the same effect is observed if the SiO_2 content of the soil is greater than 12 mg per kg soil. Similar results were found for several grains and vegetables. **W. G. R.**

Determination of easily soluble phosphoric acid and potash in soils by extraction with water and with carbonic acid-bicarbonate solutions. **B. DIRKS AND F. SCHEFFER**. *Das Superphosphat* 6, 75-80 (1930). cf *C. A.* 25, 1010 —It is proposed to det. easily sol P_2O_5 and K_2O in acid soils by extrn. with distd. water, and in neutral and alk. soils by extrn. with a H_2CO_3 soln. of $Ca(HCO_3)_2$. Thirty g. of the soil is shaken with 75 cc. of water, or the soil is mixed with 1 g. of $CaCO_3$ and shaken with 75 cc. of soln. contg. CO_2 equiv. to 80 cc. of 0.1 N NaOH per 100 cc., for 1 hr. continuously in a 100-cc. flask. P_2O_5 in the filtrate is detd. by the Mo blue method, and K_2O is detd. by the cobaltinitrite method. The method gives results which usually parallel those obtained by the Neubauer seedling method, although the actual numerical values obtained by the 2 methods may be quite different. Results obtained with the method on soils receiving the same fertilizer treatment over a period of years are tabulated, and the use or non use of P_2O_5 and K_2O fertilizers is definitely indicated by the data. **K. D. JACOB**

Increased root-solubility of potash contained in the soil as a result of phosphoric acid fertilization. **C. DREYSPRING AND F. HEINRICH**. *Superphosphate* 4, 1-10, 46-59 (1931). —As a result of expts. with the Neubauer method for root-sol. K_2O and P_2O_5 on 85 soils of different types and origin, it was found that the root-sol. of the soil K_2O was increased 1 to 14 mg. per 100 g. of oven-dry soil by the addn. of P_2O_5 fertilizers. Readily sol. phosphates were more effective in increasing the root sol. of soil K_2O than were the less sol. phosphates. Superphosphate caused an av. increase of 8.8 mg. of root-sol. K_2O and basic slag an av. of 4.2 mg. per 100 g. of oven-dry soil. The percentage increase in the assimilation of K_2O , caused by either of the 2 phosphates, systematically diminished as the root sol. K_2O content of the soil increased. The abs. assimilation by the plants, on the other hand, became slightly higher. The increase in the assimilation of K_2O was also dependent on the root-sol. P_2O_5 contained in the soil, the abs. figures became slightly, and the percentage figures considerably, smaller with increasing P_2O_5 content of the soil. D and H consider that the potash-activating power of superphosphate is due principally to its content of free H_2PO_4 and $CaH_2(PO_4)_2$. The Ca sulfate in superphosphate may also have some effect. **K. D. JACOB**

History of the ammonium citrate method for determining insoluble phosphoric acid. **W. CATESBY JONES**. *Am. Fertilizer* 73, No. 11, 36, 38, 60 (1930). —A review. **W. H. ROSS**

The fertility of the fundamental soil zones of the Union and their fertilizer requirements. **A. N. LEBEDYANTZEV**. *Udobreni i Urozhai (Fertilizers and Crops)* 2, 331-63 (1930). —The soils were analyzed chemically for the properties which indicate to some degree the fertility of the soil, such as exchange capacity, exchange acidity, hydrolytic acidity, hygroscopic moisture, N, P, K and org. matter. The results are presented graphically and in tabular form on the podzols, gray forest soils, degraded chernozem, deep chernozem, ordinary chernozem, southern chernozem, chestnut and gray and soils. The control pots of the vegetation expts., conducted by the Mitscherlich method, show that the natural fertility of the soil zones follows from a low point in the podsol zone to the highest point on the deep and ordinary chernozem and to a lower point on the southern chernozems and chestnut soils and still lower on the gray soils. The yields follow the properties of soil fertility enumerated above, which in themselves increase and decrease in the order of the soil zones as enumerated. The field expts., as well as some of the vegetation expts., indicate that all zones show deficiencies decreasing in the order N, P, K. The podzols have the least K. They also respond most effectively to fertilizers in general. The same is true for desert soils when water is supplied. Next in order of effective response are the northern chernozems. The lowest response is obtained on the

southern portion of the chernozem zone. Some specific crops require specific fertilizers irrespective of the soil zones. J S JORRIS

Survey of the fertilizer industry. P. E. HOWARD. U. S. Dept. of Agr., *Circ.* 129, 1-22 (1931).—A statistical review of the world consumption of N, P_2O_5 , and K_2O , the world production of inorg. N compds., basic slag, superphosphate and other fertilizer materials, the quantity and value of fertilizer imports and exports for the U. S.; and the U. S. supply of N, P_2O_5 , and K_2O . W H ROSS

Fertilizer trials at Rothamsted. Rothamsted Expt. Sta., *Rept.* 1929, 125 pp.—Efficiency of phosphate fertilizers on grass land is in the order superphosphate > high-sol. basic slag > low sol. basic slag > mineral phosphate. The high-sol. fertilizers increased the proportion of phosphate in the herbage. Superphosphate increased the potato crop provided no N was given, but decreased it in the presence of N and potash. In cooking tests the color of potatoes which had received potash fertilizers was superior. In this respect the chloride gave best results, followed, in order, by the sulfate and potash fertilizer salts. In flavor only the best of the potatoes fertilized with potash equalled that of the unfertilized, all others were inferior. Complete fertilizers gave best results with sugar beet, and these were still further improved by dressings of salt. With barley, H_2CN_2 and NH_4Cl again gave as good results as $(NH_4)_2SO_4$, though still inferior to $NaNO_3$. Potash reduced the attack of *Fusarium culmorum*. In pot culture chlorides delayed the ripening of barley straw and gave a lower dry-wt. yield than sulfates. The highest-quality barley was obtained with NH_4Cl and urea. No harm resulted from sowing up to 2 cwt. of H_2CN_2 per acre with the seed, but 4 cwt. per acre was distinctly injurious. B C A

Nitrogen fertilizers. BRUNO WANSKA. *Metallber.* 20, 1853 (1930), cf. C. A. 25, 1617.—A review of various pyrolytic and acid processes for making mixed fertilizers from phosphate rock and K Al silicates, from phosphate rock, HNO_3 , and NH_3 , and from KCl , H_3PO_4 , and NH_3 . S L MADORSKY

Results of experiments with nitrogen fertilizers on cotton. V. N. MIKHAILENKO. *Udobreñi i Urozhai (Fertilizers and Crops)* 2, 363-9 (1930).—The addition of nitrogenous fertilizers from various sources at the time of planting shows that the best results during the year of application were from $(NH_4)_2SO_4$, followed by KNO_3 , $CaCN_2$, NH_4NO_3 , and $CO(NH_2)_2$. The best after-effects were shown by NH_4NO_3 , followed by $(NH_4)_2SO_4$, KNO_3 , $CO(NH_2)_2$, and $CaCN_2$. The highest total increase in yield for the 2 years came from $(NH_4)_2SO_4$, followed by NH_4NO_3 , KNO_3 , $CaCN_2$, and $CO(NH_2)_2$. When the same fertilizers were added during the process of boll formation the $CaCN_2$ was the most effective during the first and second years. In general, the increase in yield was higher when the fertilizers were added later in the process of growth. This is in accordance with the results of physiologist experiments which show that cotton takes up more N in the later stages of growth than during the early part. J S JORRIS

Nitrate fertilization and nitrogen fixation by the legumes. LAWIN WEBER. *Zentralblatt Parasitenk.* 2 Abt., 82, 353-79 (1930).—Nitrate fertilization decreases N fixation. JOHN T. MYERS

Potash fertilizers at the "Belle Etoile" distillery. DUBOSC. *Bull. agr. Congo Belge* 21, 581-3 (1930).—The addition of a large quantity of K_2O causes a richer and purer juice in the sugar cane and favors the maturity of the cane. This in turn favors a greater production of rum to the acre of cane. J R A

Tests with phosphate fertilizers. G. TOMMASI AND S. DOJMI. *Atti III congresso nat. chim. pura applicata* 1930, 561-3.—In some cases the use of finely ground mineral phosphate is advantageous, but much less effective than sol. phosphate fertilizers. E M. SYMONS

The use and action of phosphates on chernozem. M. A. EGOV. *Pochvovedenie* (Russian) 23, No. 3-4, 5-28 (1928), cf. C. A. 24, 2534.—In an attempt to establish the conditions for the reversion of phosphates in chernozem soils, E. conducted a series of experiments assuming beforehand that the Ca humates in the soil undergo double decomposition with sol. phosphates. By using the mono-, di-, or tri-Na phosphate on soils the corresponding Ca phosphates ought to be obtained. Ten g. of soil was mixed with 50 cc. of 0.25 N solns. of various phosphates, shaken for 3 min. and filtered. With KH_2PO_4 , the ext. was almost colorless; with NaH_2PO_4 , the ext. was light yellowish brown; with $(NH_4)_2HPO_4$, K_2HPO_4 , and Na_2HPO_4 , it was the same; with K_3PO_4 , it was dark brown; with Na_3PO_4 , it was intense dark brown. Thus the color depends on the nature of the cation (Na being more active) and the anion ($H_2PO_4^-$ and HPO_4^{2-} giving a slight coloring and PO_4^{3-} producing an intense coloration). The intensity of the color increased with the increase in concn. and an increase in the time of shaking. The humus substances can be released by washing with a salt soln. and then with water. Ertz

obtained by the use of NaH_2PO_4 , Na_2HPO_4 , or KH_2PO_4 together with CaCO_3 are darker than those obtained with the phosphate alone. CaCO_3 alone when mixed with soil and water gives no humus in the ext. Five-g samples of soil were treated with 25 cc. 1.0 N NaH_2PO_4 , KH_2PO_4 , and Na_2HPO_4 , shaken for 3 min., placed in collodion bags and dialyzed. After a few days a brownish liquid began to rise from the surface of the settled soil in the bag with the NaH_2PO_4 and diffuse through the supernatant liquid. It finally turned the entire bag into a dark-colored mass, some org matter started to go through the dialyzer. This action was slower with the Na_2HPO_4 . With the KH_2PO_4 the liquid after 3 months was just a weak brownish yellow. The behavior of the NaH_2PO_4 and Na_2HPO_4 is explained on the basis of formation of the more sol $\text{Ca}(\text{H}_2\text{PO}_4)_2$ and the less sol CaHPO_4 . Addns of FeSO_4 or FeCl_3 to exts of Na_2HPO_4 had no effect on the mobility of the humus. Addns of FeCO_3 gave the same results as addns of CaCO_3 . The reactions in all cases resolve themselves into the formation either of a sol or an insol phosphate. By successive extns with Na_2PO_4 and decanting it was found possible to obtain in soln most of the org fraction of the absorbing complex, about $\frac{1}{2}$ to $\frac{1}{3}$ of the total org matter is contained in this fraction. J. S. JOFFE

Fertilizing action of bone meal. E. TRUNINGER AND F. KELLER. *Landw. Jahrb. Schweiz* 43, 931-45 (1929).—Bone meal, with preference for the defatted, is recommended to give good results when used on acid soil. One application of raw bone meal gave 0.33 the increase of superphosphate, on doubling the application the increase was 0.50. There is a relationship between the particle size and the activity. N from bone meal is estd. to be 50% as valuable as nitrate N. GEORGE R. GREENBANK

Ammonia liquor as a fertilizer. BRISCHAS. *Gas u. Wasserfach* 74, 13-6, *Gas Journal* 193, 324 (1931).—Pot tests carried out at the Mitscherlich Expt. Sta. at Marienburg on the effect of NH_3 liquor on oats and mustard plants in light, medium and heavy soils indicate that good results may be obtained on the better or heavier soils. Phosphate and potash were added in each case. In spite of the alk. nature of the NH_3 , better results were obtained when phosphate was added as Thomas meal instead of superphosphate. This was first noted at a period about half way in the growth of the plants and became more noticeable as the expt progressed. Poor results were obtained on a pure sand soil, even though potash and phosphate were added. R. W. RYAN

Composition of lime for fertilizer. YOSHIZO HAYASHI. *J. Imp. Agr. Expt. Sta.* 1, 57-72 (1929).—The CaO content of 43 representative liming materials in Japan varied from 55.94 to 95.92%, av. 71.82%. The finer the particles of lime, the lower was the CaO content. K. KITSUTA

Production of citric-soluble fertilizers by the decomposition of phosphoric with silica at high temperatures. JAAK KUUSK. *Acta Commentationes Univ. Tartuensis (Dorpatensis)* AXI, 1, 1-79 (1930) (in German).—A study was made of the optimum conditions for the formation of citric sol phosphate (sol in 2% citric acid soln) by heating natural phosphates with SiO_2 . With Esthonian phosphorite and Florida phosphate, both of which have the fluorapatite structure, active decompn began at about 1370° and, in general, the percentage of citric sol P_2O_5 increased with further increase in temp. With finely ground phosphate, passing a sieve having 5000 meshes/sq cm, the amt of citric sol P_2O_5 formed depended upon the quantity of SiO_2 used, up to about 42% of the wt of the furnace charge, and better decompn was obtained with finely ground SiO_2 , passing a sieve having 5000 meshes/sq cm, than with the same wt of coarse SiO_2 . Approx. 90 to 98% of the total P_2O_5 was converted into the citric sol form when mixts contg 30 to 40% SiO_2 were heated rapidly in an internally fired gas furnace to 1380° and then heated to 1440-1460° at the rate of 2°/min. Lower results were obtained when the mixts were heated in an elec. furnace under the same conditions. Phosphates contg excessive amts of lime were successfully decompd by proportionately increasing the amt of SiO_2 in the charge. The presence of more than about 10% of Fe_2O_3 and Al_2O_3 depressed the citric soly of the P_2O_5 . K. thinks that the citric sol P_2O_5 is present as a *silicophosphate*, $3\text{Ca}_3(\text{PO}_4)_2 \cdot \text{CaSiO}_3$, which results from the replacement by SiO_2 of the F in the fluorapatite compd, the F being volatilized as SiF_4 . With phosphates of the carbonate-apatite type decompn began at about 1100°, and above this temp. the percentage conversion in a given time increased rapidly.

Industrial value of marine flora. VICTOR DE BUEN. *Rev. Ing. ind.* 1, No. 6, 5-12, No. 8, 12-6 (1930).—A brief review is given of methods of prepn of fertilizers, I, alginate acid, cellulose, food products and other org preps from various marine algae and phanerogams. L. E. GILSON

Field results indicate value of ammoniated superphosphate for cotton. ANON. *Phosphorus Digest*, Feb., 1931, 2-6.—Complete mixts prepd from KCl and highly

ammoniated superphosphate plus $(\text{NH}_4)_2\text{SO}_4$ (when additional N was required) were compared with mixts. of similar analysis prepd from KCl, ordinary superphosphate and $(\text{NH}_4)_2\text{SO}_4$ as fertilizers for cotton in field tests carried out in S. C., Ga. and Ala. In 17 tests the ammoniated superphosphate mixts gave an av. yield of seed cotton 64% greater than that obtained with the ordinary superphosphate mixts. The av. yield of cotton obtained at the first picking was also higher from the plots fertilized with the ammoniated superphosphate mixts. The ammoniated superphosphate used in these expts was prepd by treating ordinary superphosphate with approx. 6% of anhyd. NH_3 .

K. D. JACOB

Effect of fertilizers on the phosphorus content of plants. J. H. MITCHELL. *Phosphorus Digest* March, 1931, 6-7—1 expts. by the S. C. Expt. Sta. over a no. of years showed that application of complete fertilizers contg. 10-12% P_2O_5 gave av. increases of 27, 42 and 47% in the P_2O_5 content of oats, soy beans and grass hay, resp.

K. D. JACOB

The present and the improved methods of making farmyard manure in Kanara. R. M. HEGDEKATTI. *Poona Agr. Coll. Mag.* 21, 222-40 (1940)—Large quantities of farmyard manure are prepd. in India by the fermentation in pits of org. materials in the presence of approx. 30% by wt. of cattle dung and urine. The principal org. materials used are the green and dry leaves and young shoots of forest plants. The leaves of a no. of varieties of plants used for this purpose contain approx. 1.10 to 2.10% N and 0.5 to 2.0% tannin as well as astringent resins and volatile oils which act as deterrents to injurious insects in the soil. Fermentation of the material in the pits is hastened by keeping it moist and turning it occasionally. Fermentation is also hastened by the addn. of 0.1% N in the form of $(\text{NH}_4)_2\text{SO}_4$ and 5% of lime or lime alone. A better grade of manure is produced in covered pits than in uncovered pits. In 7 expts the loss of org. matter during fermentation ranged from 29.5 to 64%. Analyses of various org. materials and of manures prepd. from them are tabulated.

K. D. JACOB

Barnyard manure nitrification in cultivated fields. CHAS. BARTHEL AND N. BENGTSSON. *Agr. Landbruks Akad. Handl. Tid.* 69, 1014-31 (1940)—1 expt. with pure bacterial, yeast and mold substance and different amino acids, show that N which is included as a constituent in the cell substance of different microorganisms, and which forms the main portion of the org. N in barnyard manure, is nitrified readily in the soil when it is present in the form of pure preps. However, org. N in barnyard manure does not nitrify in tilled soil until a very long time has elapsed, and has no fertilizing effect during the first months or years after introduction. Either the microbial cells in the manure are protected from decompn. by colloidal substances in which they are imbedded or the decompn. depends upon the C-N ratio (cf. Jensen, *C. A.* 23, 5243). Data are insufficient to decide upon either theory.

E. M. SUMMERS

The conservation of ammonia in fertilizing with liquid manure. S. TOYBERG-JENSEN. *Uss Arch. Landw.*, Abt. A. *Pflanze* 3, 161-89 (1930)—The rate of escape of NH_3 from solns. of $(\text{NH}_4)_2\text{SO}_4$ plus NaOH and NH_4HCO_3 was detd. by drawing air through the solns. maintained at const. temp. Lab. expts. were conducted on the rate of escape of NH_3 from liquid manure spread on various soils. The effect of added salts was detd. by adding CaCl_2 , KNO_3 , gypsum and superphosphates to the liquid manure, and detg. the NH_3 lost when the solns. were allowed to dry on filter papers. The lab. expts. show that the liquid manure is best spread on the soil when it is porous in cool weather and that addn. of the salts mentioned prevents 90% of the loss of N_2 ordinarily suffered. Field expts. were also conducted by fertilizing with (1) liquid manure, (2) KNO_3 , CaCl_2 and superphosphate dissolved in liquid manure and (3) liquid manure and each of these salts added separately. The best results were obtained by dissolving the salts in the manure. The following rapid volumetric method for the detn. of NH_3 is described. The liquid manure is added to an alk. NaOBr soln. and the N_2 evolved is collected and measured in a gas buret. 100% of the N_2 in NH_3 is obtained by this method, but only 95% of the N_2 is liberated from urea.

W. GORDON ROSE

The use of preservatives to prevent loss of nitrogen from cow excreta during the day of collection. ROWLAND B. FAENCH. *J. Agr. Research* 41, 503-6 (1930)—Thorough mixing of the feces and urine greatly increases the speed of the action by which N is lost. When unmixed this loss is from 1 to 2.6% of the total N in 18 hrs. at 70° F. and when these excreta are thoroughly mixed it varies from 6 to 30%. CuSO_4 , ZnSO_4 , and formaldehyde in small concns. gave complete protection against loss of N when added to homogeneous mixts. of feces and urine but the CuSO_4 was very ineffective with unmixed excreta. At higher concns. boric acid, FeSO_4 , phenol, salicylic acid and toluene also showed some preservative action.

J. R. ADAMS

The chemical composition of grass from plots fertilized and grazed intensively in

1929. J. G. ARCHIBALD *J Agr Research* 41, 491-501 (1930).—The dry matter of the grass from the fertilized plots was higher in N, ether ext., Ca and P, and lower in crude fiber than the dry matter of the grass from the unfertilized plot. Acre production of all constituents detd., crude fiber excepted, was markedly higher on the fertilized than on the check plots. The seasonal variations in the compn. of the grass confirm the conclusions of a previous paper (cf. *C A* 23, 4008). The amt. of rainfall is one of the most important factors which influence the chem. compn. of the grass. J. R. ADAMS

Club-root in cruciferous crops. J. G. GRASS. *New Zealand J Agr* 42, 1-17 (1931).—Fertilization of cruciferous plants with basic slag and K_2SO_4 seemed to reduce the intensity of club-root infection. Satisfactory control of the disease was obtained by applying burnt lime at the rate of 2 to 3 tons per acre about 3 months before sowing the crop. K. D. JACOB

Calcium relationships of forage crops. L. I. PUGSLEY AND R. R. MCKIBBIN. *Can J Research* 4, 39-51 (1931).—Samples of soil, timothy and red clover were secured from the Chateaugay River basin, Quebec, and analyzed to det. the cause of the variation in Ca content of clover and timothy previously observed. The Ca content of timothy and red clover increased directly with the total Ca content of the underlying soil and decreased as the 'lime requirement' of the soils increased. No proportionate increase of protein or of P as the Ca content of the hays increased was observed. The timothy hays from the area investigated were extremely low in Ca content but otherwise normal.

The ratios $\frac{CaO}{P_2O_5}$ and $\frac{protein}{CaO}$ for these hays are, resp., abnormally low and abnormally high. In the red-clover hays these values were normal. J. W. SHIPLEY

Nature and course of the nutrient intake of various vegetables. T. REMY AND F. WEISSE. *Lands Jahrb* 71, 315-331 (1930).—The rate of intake of a no. of vegetables is studied by cultural trials. The significance of the cultural intake and growing period is considered in studying the proper fertilizer treatment and the appropriate fertilizer. GEORGE R. GREENBANK

Fertilization of cucumbers and strawberries with carbon dioxide gas. H. LUNDEGÅRDH. *Kgl Lärskrugs Akad Handl Tid* 69, 1044-56 (1930).—The theoretical basis of CO_2 fertilization has been investigated (Svenges Trädgårdsodling Årshok 1923), and practical tests with CO_2 fertilization on various hot house plants have been made (Der Kreislauf der Kohlensäure in der Natur, Jena, 1924, Centralanst. f. försöksväs. på jordbruket, Medd. no. 331 (1928)). Tests were repeated in hot houses, with alc. burned as the source of CO_2 . A l. of 95% alc. gives 6000 cal. on burning and induces air currents for distribution. CO_2 fertilization increases the fruit no. and wt. of cucumbers 25.5-45%, with a CO_2 atm. 3-4 times normal. Lamps were filled every morning, lighted, and burned about 20 hrs. until extinguished spontaneously, then the dil. alc. residue was removed to avoid evapn. of denaturant. With strawberries, ripening was earlier, the fruit no. increased 12%, the wt. of fruit higher, and the leaf growth greater. A profit is shown by comparison of the alc. purchased vs. increased return from sale of cucumbers. E. M. SYMMES

Potato scab and fertilizers. EICHINGER. *Das Superphosphat* 7, 8-14 (1931).—Superphosphate and $(NH_4)_2SO_4$ have a decided effect in reducing scab on potatoes. Sulfate of potash magnesia (a mixt. of K_2SO_4 and $MgSO_4$) and probably K_2SO_4 tend to reduce the occurrence of scab. KCl has no protective action, but seems to favor scab formation to a slight extent. Because of its alk. nature basic slag favors scab formation. Scab formation is promoted to a marked extent by $NaNO_3$ and to a less extent by $Ca(NO_3)_2$. K. D. JACOB

Preparations containing rotenone for use as insecticides. L. Aqueous suspensions. HOWARD A. JONES AND WILLIAM M. DAVIDSON. *J Econ Entomol* 24, 244-57 (1931).—Aq. suspensions of rotenone were prepd. by two methods: (1) Rotenone solns. in water miscible solvents were added to water, and powdered preps. were mixed with water. (2) Permanent suspensions contg. 0.05-0.5 g. rotenone per l. with the rotenone in highly dispersed condition were prepd. by the addn. to water of pyridine solns. and acetone solns. contg. tannic acid. Mixts. of certain dry products with pptd. rotenone also gave satisfactory suspensions. Permanent suspensions of rotenone were more easily obtained under slightly alk. conditions, but they were difficult to obtain in hard water. Toxicity expts. on various species of insects indicated that aq. suspensions of rotenone are of definite value in the control of certain types of insects. Rotenone decomposes rapidly in pyridine, less rapidly in acetone contg. tannic acid and very slowly in pure acetone soln. This decompn. is accompanied by a corresponding loss in toxicity (cf. following abstract). Pure dry rotenone shows no change in toxicity on standing. C. H. RICHARDSON

The toxicity of rotenone in solution and suspension. W. M. DAVIDSON AND HOWARD A. JONES. *J. Econ. Entomol.* 24, 257-62 (1931).—The loss in toxicity to insects of rotenone in pyridine solns. and acetone solns. contg. tannic acid is described. The pure acetone and alc. solns. of rotenone did not lose toxic activity upon standing. The yellow decompn. material obtained in a pyridine soln. of rotenone was much less toxic to insects and fish than pure rotenone. Aq. suspensions of rotenone made from fresh acetone and alc. stock solns. decreased in toxicity to insects upon standing. It is recommended that stock solns. of rotenone in pyridine and acetone be used immediately after prepn. Solns. in acetone and alc. may be allowed to stand a limited time without loss in toxicity. C. H. RICHARDSON

Some comparative tests with rotenone, nicotine and pyrethrum. MERRILL M. DARLEY. *J. Econ. Entomol.* 24, 111-5 (1931).—A soln. of rotenone (concn. 1:100,000) contg. a com. petroleum prepn. (concn. 1:200) compared favorably as a contact poison with nicotine at 1:10,000 and pyrethrum at 1:74,800 concn. the tests being made upon 2 species of aphids. Pyrethrins at 1:37,400 were effective against the spotted cucumber beetle (*Diabrotica duodecimpunctata*) and the Mexican bean beetle (*Epilachna corrupta*), but nicotine at 1:5000 and rotenone at 1:50,000 were ineffective. A com. petroleum prepn. at 1:100 was used in each spray. Tests indicated that the toxicity of rotenone for the common red spider (*Tetranychus biatus*) was appreciably greater than that of nicotine and pyrethrins. The relative toxicity of rotenone was greater for this animal than for the aphids. C. H. RICHARDSON

Genuine dennis root may contain no rotenone. R. C. ROARK. *J. Econ. Entomol.* 24, 328-30 (1931).—Expts. have shown that rotenone is far superior in insecticidal value to deguelin, tephrosin and toxicarol, the other insecticidal constituents of dennis. Therefore all the matter extd. with ether from dennis root may not be equal to rotenone in toxicity. Numerous samples of root sold as dennis root have recently been found to contain no rotenone. The rotenone content of dennis root varies from 0 to 5.5%, whereas the total ether ext. ranges from 5 to 23%. There is no consistent relation between the two. The compn. of the ext. from dennis root, *s. e.*, the ratio of one constituent to the other, also may vary considerably. A sample found to be rich in toxicarol entirely lacked rotenone. This lack of uniformity in proprietary dennis preps. accounts for many of the inconsistent results reported by entomologists. The common dennis root now on the market may be expected to contain about 2% rotenone. Cubé root (*Lonchocarpus wron*) from South America contains as much as 7.6% rotenone. An ext. of cubé root prepd. in Peru or Brazil may soon be placed on the American market. C. H. RICHARDSON

An insecticidal method for the estimation of kerosene extracts of pyrethrins. H. H. RICHARDSON. *J. Econ. Entomol.* 24, 97-105 (1931).—The speed with which pyrethrum exts. paralyze house flies (*Musca domestica*) varies directly with the concn. The speed of paralytic action is variable, but the variation falls within the limits of error due to random sampling. This was not true of the mortality after 24 hrs. Differences in speed of paralytic action between a 0.75 lb. (per gal.) and 1 lb. kerosene ext. of a fairly powerful grade of pyrethrum powder (0.26% pyrethrin I) were easily distinguishable and statistically significant. Differences in percentage mortality, analyzed similarly, were not significant. C. H. RICHARDSON

Problems in the manufacture of liquid household insecticides of the petroleum extract of pyrethrum type. ALFRED WHEED. *J. Econ. Entomol.* 24, 95-7 (1931).—The problems which confront the manufacturer of pyrethrum products are: (1) the concn. of the toxic constituents to be used, (2) the use of pyrethrum powder or cone. exts. in prepg. the products, (3) the selection of satisfactory petroleum fractions and perfumes. C. H. RICHARDSON

The effect of soap on the toxicity of a pyrethrum product known as "Red Arrow." A. E. BADERFSCHER. *J. Econ. Entomol.* 24, 268-77 (1931).—This is a study of a com. mixt. contg. pyrethrum ext. with the addn. of K. coconut fatty acid soap in various proportions. The green spruce aphid (*Aphis spruceola*) was used as the test insect. "Red Arrow" did with tap water does not give max. efficiency probably because the surface tension of the aq. carrier is too high to permit sufficient penetration into the trachea of the insect. The toxicity of this product is increased about 4 times when did. with 0.4% soap, the surface tension of the aq. carrier then being nearly 30 dynes. "Red Arrow" loses little if any of its toxicity upon standing. C. H. RICHARDSON

Experiments with white oil-pyrethrum for the control of the oriental peach moth. BYRLEY F. DRIGGERS. *J. Econ. Entomol.* 24, 319-25 (1931), cf. C. A. 24, 3313.—The newly hatched larvae of *Laspeyresia molesta* are killed more easily when crawling over sprayed new growth than when crawling over sprayed old growth, because of the great

absorptive power of the old growth for the oil. Reduction of fruit infestation amounted to 30-70%. The effects of parasites and their relation to spraying operations are discussed. C. H. RICHARDSON

Tests with nicotine activators. E. P. FELT AND S. W. BROMLEY. *J. Econ. Entomol.* 24, 105-111(1931).—Field tests were carried out with nicotine solns. contg. various petroleum oil emulsions, Na oleate, K oleate, com. soap flakes and Ca caseinate preps. The results indicate the influence of a complex of factors on the toxicity of the various mixts. C. H. RICHARDSON

Nicotine in the control of ectoparasites of poultry. T. C. BISHOPP AND R. D. WAGNER. *J. Econ. Entomol.* 24, 56-62(1931).—The feather mite (*Liponyssus sulzarium*) is controlled by treating roosts and spraying the houses with nicotine sulfate soln. The common chicken mite (*Dermanyssus gallinae*), the fowl tick (*Argas miniatus*) and the sticktight flea (*Echidnophaga gallinae*) are controlled by spraying poultry houses with strong solns. of nicotine sulfate. Fowls are treated for the various species of chicken lice by making a light application of nicotine sulfate (40%) to the upper surface of the perches $\frac{1}{2}$ hr. before the fowls go to roost. With proper ventilation of poultry houses and correct dosage this method will not injure fowls. Disadvantages of the method are discussed. C. H. RICHARDSON

Comparative performance of nicotine tannate and lead arsenate against the codling moth. ROBERT S. FILMER. *J. Econ. Entomol.* 24, 277-83(1931).—If a sufficient quantity is maintained upon the foliage during the period of codling moth activity nicotine tannate controls this insect as well as Pb arsenate. Nicotine tannate in appropriate concns. injures neither fruit nor foliage. In its present form nicotine tannate disappears rapidly from the foliage, 60-70% loss occurring in 10 days after application. Before it can be used as an efficient substitute for arsenicals its adhesive qualities must be improved. C. H. RICHARDSON

Report of committee to formulate plans for investigation of the codling moth from biologic and control standpoints. A. L. QUAINANCE, LEROY CHILDS, G. A. DEAN, D. A. PORTER, P. J. PARROTT AND W. A. ROSS. *J. Econ. Entomol.* 24, 18-23(1931).—Results of lab. and field tests from 23 Federal and State entomological labs. are given. The following insecticides are considered: Mn arsenate, Na_2SiF_6 , BaSiF_6 , K fluoaluminate, petroleum oil-nicotine mixt., petroleum oil-pyrethrum mixt., petroleum oil-Pb arsenate mixt., petroleum oil-derris mixt., highly refined (white) petroleum oil emulsion, nicotine tannate and rotenone. Removal of As spray residue also is considered. Mech. control methods (banding, chemically treated bands, bait traps, soil treatment), use of parasites and biol. data are reported. C. H. RICHARDSON

The effect of hydrogen-ion concentration on the toxicity of nicotine, pyridine and methylpyrrolidine to mosquito larvae. CHARLES H. RICHARDSON AND HAROLD H. SHEPARD. *J. Agr. Research* 41, 337-49(1930).—The toxicity of 0.03 M nicotine solns. to mosquitos increased with increasing pH value, reaching a max. at pH 9.7 (the free base). Similar results were obtained when either HCl or H_2SO_4 was used to adjust the pH of the soln. At each nicotine concn. the free base was about 5 to 7 times more toxic than was nicotine sulfate soln. at pH 5.0. The addn. of an inorg. hydroxide had no apparent effect upon the toxicity of a nicotine soln. Pyridine solns. of 0.03 M and 0.12 M were about twice as toxic as solns. of the same molar concn. titrated to pH 5.0 and 4.9, resp., with H_2SO_4 . Methylpyrrolidine soln. was about 19 times as toxic as methylpyrrolidine-HCl soln. of pH 3.0. The toxicity of these materials in an soln. is directly related to the concn. of the undissocd. mols. The toxicity of nicotine solns. is largely dependent on the disson. of the pyrrolidine N. J. R. ADAMS

Trapping for the Japanese beetle (*Popillia japonica* Newman) during the seasons of 1929 and 1930. O. K. COURTNEY. *J. Econ. Entomol.* 24, 219-26(1931).—The operation of traps with a bait mixt. contg. geraniol and eugenol is described. C. H. RICHARDSON

The tolerance of beans to sprays and dusts for the Mexican bean beetle. H. C. HUCKERT. *J. Econ. Entomol.* 24, 200-4(1931).—Further studies with com. brands of arsenicals, BaSiF_6 , and cryolite in spray and dust mixts. against *Epilachna corrupta* indicate that certain forms of arsenicals cannot always be used with safety upon bean plants. Mg arsenate, BaSiF_6 , and cryolite are the safest to use at recommended concns. C. H. RICHARDSON

Hints at new spraying procedures. P. J. PARROTT, F. Z. HARTZELL, HUGH GLASGOW AND S. W. HARMAN. *J. Econ. Entomol.* 24, 297-302(1931); cf. C. A. 24, 3312.—This investigation was conducted to find materials or practices that would amply present control measures for the rosy aphid (*Anuraphis roseus*), and the eye-spotted bud moth (*Spilonota ocellana*) on apple trees. NaNO_3 , one lb. per gal., made the trees unattractive

to rose aphid without apparent injury to the trees. Fall applications of nicotine or ground tobacco to apple trees reduced the number of aphids. Tar distillate sprays were very effective against apple aphid eggs and nymphs of the spruce gall aphid (*Adelges abietis*) but were ineffective against eggs of the leaf roller. These tar distillate sprays appeared safe to use on the trees. American crocote oil emulsified with sulfonated castor oil and alkali killed eggs of the rose aphid and caused no serious injury to the trees. Sulfonated fish oils apparently may be used in place of the sulfonated castor oil to reduce the cost of this spray. These methods may offer a marked advantage over spring spraying for control of the rose aphid. C. H. RICHARDSON

A study of arsenical residue on apples in Pennsylvania with respect to efficient spraying practice. H. E. HODGKINS AND D. E. HALEY. *J. Econ. Entomol.* 24, 80-8 (1931).—In Pa. protection of apples from second brood larvae of the codling moth (*Carpocapsa pomonella*) requires a definite spray procedure ending early in June. In this way it is possible to control the insect with Pb arsenate and to keep the As residue within the export tolerance of 0.1 gram per lb. of apples. Analyses of fruit are given. C. H. RICHARDSON

What summer oil sprays may do to apple trees. JOSEPH M. GIVERTZ. *J. Econ. Entomol.* 24, 283-90 (1931).—Refined (white) petroleum oils, viscosity 80-120° Saybolt at 100°F., and sperm oil were sprayed on apple and peach trees during the spring and summer months. The oils were emulsified with soap or powdered skim milk, and the emulsions dil'd. to contain 0.5-3% actual oil. Some trees were sprayed during two successive seasons and accumulative effects on the fruit and foliage were studied. Peach trees are less susceptible to petroleum oil injury than apple trees. Continuous oil spraying of apple trees during the growing season produced physiol. changes in growth processes and fruit production. Oil injury is more pronounced and more severe from early than from late spraying. Sperm oil was less injurious to peach and apple foliage than any of the other animal or vegetable oils tested. C. H. RICHARDSON

Experiments in spray residue removal. L. R. STREETER, P. J. CHAPMAN AND G. W. PRACE. *J. Econ. Entomol.* 24, 240-4 (1931).—The standard washing machine using HCl as a solvent entirely and satisfactorily removes As residue. Methods successful in the Pacific Northwest are not entirely adaptable to the needs of the Northeast. In the Northeast a machine is required which combines the washing and wiping principle. No single machine now available entirely meets these requirements. Results of chem. analysis of sprayed fruit are given. C. H. RICHARDSON

Investigations aimed to reduce the cost of pear psylla control. F. Z. HARTZELL. *J. Econ. Entomol.* 24, 71-7 (1931).—Economy can be obtained in part by modifying the environment of the orchard, by using lubricating oil emulsions, and by decreasing the concn. of nicotine in the sprays. Care must be exercised in the use of petroleum oil sprays on weak trees. C. H. RICHARDSON

The compatibility of lime with fluosilicates. R. H. CARTER. *J. Econ. Entomol.* 24, 263-8 (1931), cf. *C. A.* 24, 455.—The compn. and chem. properties of fluosilicates in relation to insecticidal use and injury to foliage are described. The addn. of small quantities of Ca(OH)₂ to Na₂SiF₆, K₂SiF₆, and BaSiF₆ materially decreased the solv. probably because of the introduction of a common ion from the reactions. Strong alk. solns. were formed when an excess of lime was present as the result of further chem. reactions. C. H. RICHARDSON

Mercury salts as disinfectants of seeds. VITTORIO CASABURI. *Industria chimica* 5, 1362-8 (1930), cf. *C. A.* 25, 553.—By using an emulsion of 1.4-Cl₂C₂H₄ in petroleum taken up with talc in conjunction with Hg salts, a powerful disinfecting action is obtained, with a smaller quantity of Hg salts. A. W. CONTEHR

Tests with *p*-dichlorobenzene-oil emulsion against the San José scale. OLIVER I. SNAPP AND J. R. THOMSON. *J. Econ. Entomol.* 24, 328 (1931).—A petroleum oil of 168 sec. viscosity (Saybolt) and a volatility of 0.60% was used in these expts. *p*-C₆H₄Cl₂ was dissolved in it in the proportion of 2 lb. per gal. The oil was emulsified with K fish oil soap and the spray contained 0.25-1.5% actual oil. The addn. of *p*-C₆H₄Cl₂ to an oil does not increase its effectiveness for the control of San José scale. C. H. R.

The influence of plant growth on the activity of root bacteria (G&S) 11C. Removal of spray residues from apples (NELLER) 12. Burning accident from NaClO₂ (DEEM) 13. Effect of various Mn compounds on the quinhydrone electrode (OSCOL KASHIHARA) 2. "Lime-nitrogen" (Brit. pat. 326 677) 18. Apparatus for making lime-nitrogen (Ger. pat. 516,445) 18. Dicalcium phosphite and KNO₃ (Brit. pat. 326 692) 18. Purification of gases and the manufacture of fertilizers (Belg. pat. 371,622)

Fertilizer. C. VIRAIN. Brit 330,651, July 15, 1929. Natural dolomite is treated with H_2PO_4 and the product may be treated with NH_3 .

Fertilizers. SOC D'ETUDES POUR LA FABRICATION ET L'EMPLOI DES ENGRAIS CHIM. Fr 601,155, July 27, 1929. A compd fertilizer is prepd by treating natural chlorides with a strong mineral acid such as H_2SO_4 , HNO_3 or H_2PO_4 , isolating the K salt formed and using the liberated HCl to attack natural phosphates and mixing the sol phosphate with the K salt. Cl C. A. 25, 1625.

Fertilizers. SOC D'ETUDES SCIENTIFIQUES ET D'ENTREPRISES INDUSTRIELLES. Fr. 601,628, April 28, 1930. A fertilizer is composed of a mixt. of natural phosphates and blast furnace or P. furnace slugs.

Fertilizers. G. OBER & SONS CO. Fr 601,113, April 17, 1930. Phosphatic fertilizers are made by sending a current of hot acid through a restricted receiving zone, mixing this current with a current of finely divided phosphatic material and sending the mixed current through the restricted receiving zone. The phosphatic powder may be aerated to render it more fluid. An app. is described. Fr 601,111, describes the prepn. of a fertilizer by submitting a phosphatic material to the action of an acid in a closed vessel and afterward introducing NH_3 into the vessel. Fr 601,116, describes the prepn. of a fertilizer by forming currents of acid and finely divided phosphatic material, mixing the currents, introducing the mixed current into a vessel and regulating the concn. of the acid current and the duration of the charging to reduce to a min. the losses of volatile substances during the transport.

Powdered deodorizer for accelerating the ripening of organic fertilizers. KAZU-YOSHI YAMAJI. U. S. 1,701,018, Feb. 10. A cereal powder contg. a large quantity of enzymic material such as barley, wheat or beans is mixed with bran, water is added and the mixt. is heaped for fermentation, after which it is dried and powdered and mixed with dried and powdered germinated seeds of cereal such as barley, wheat or beans high in enzymic content and tri-Ca phosphate also is added to the mixt., which, as thus prepd., is suitable for treating org. fertilizer or use as a deodorant.

Insecticides. MAURICE BOURMESTONNET. Fr 601,130, April 18, 1930. An insecticide and rat-killer is composed of a mixt. of HCN with 3-6% of oxalic acid and 2-5% of CNBr .

Insectifuges. DRUG-CHEN GOLD-UND SILBER-SCHIEDERSTALT VOM ROSENLEIN. Ger. 510,373, Apr. 20, 1924. Pulverulent compns. for combating flea-beetles, etc., comprise a solid absorbent, e. g., Kleieguhr, mixed with nitrobenzene and a substance which repels the insects by its odor, e. g., "oleum animale," phenol or cresol. Substances improving the adhesion, e. g., talc, may be included.

Destroying insects and other animal pests. I. G. FARRINGTON A. G. (Carl Taube and Hans Kükenthal, inventors). Ger. 510,406, Oct. 10, 1929. Esters of H_2SO_4 are used, alone or mixed with diluents, other insecticides, inflammability-reducing substances, etc.

Bactericides; fungicides. I. G. FARRINGTON A. G. (Rudolf Lehmann and Richard Michel, inventors). Ger. 510,485, Nov. 21, 1924. Mixts. of halophenols (or their derivs.) with esters of substituted aromatic carboxylic acids are used, with or without other disinfectants, etc. A mixt. of trichlorophenol acetate 1 and methyl *p*-hydroxybenzoate 3 parts is specified.

Plant-protecting agents. I. G. FARRINGTON A. G. Ger. 513,776, Dec. 6, 1927. The agent consists of hydrocarbons of the diaryl type, combined with other fungicides or insecticides and optionally mixed with solvent, adhesive or spreading agents. The examples mention biphenyl, 1,1'-dichlorobiphenyl, 2,2'-diaminobiphenyl and dinitro-biphenyl.

Mordants for seeds. I. G. FARRINGTON A. G. Fr 601,001, May 6, 1930. Mordants for seeds are composed of thio-derivs. of alkylarsenous acids such as methylarsine sulfide, or the condensation product of methylarsine and β -thionaphthol, arsenomethane and diethyglycolic acid or methylarsine oxide and thioacetic acid. Cl C. A. 25, 1620.

Herbicide. GUSTAV LITSCHE. Pr 601,663, April 20, 1930. A herbicide is composed of CuSO_4 and about 3 times the amt. of H_2SO_4 of 60° B.

Weed killers. CHEM. FAN, OKTA & BRAUNSCHEIDT A. G. Ger. 510,141, Dec. 6, 1924. The corrosive action on metal tanks, sprayers etc., of weed killers comprising heavy metal salts, is reduced by addn. of a nitrate, sulfate or thiosulfate of an alkali, alk. earth, or earth metal. The addn. must not give an insol. ppt. with the heavy metal salt. Thus, 1% of NaNO_3 may be added to a 2% soln. of CuCl_2 .

16—THE FERMENTATION INDUSTRIES.

C. N. FEET

Constitution of fusel oil obtained from fruits. G. VECCHI AND P. HALLER. *Mitt. Lebensmittel Hyg.* 21, 321-5 (1930) (in Italian).—The constitution of fusel oil depends upon the source. An example is a limpid, yellow brown liquid, d_4^{20} 0.800-0.800, acid trace, esters 13% by vol. calcd. as EtOAc , furfural 0.01, aldehydes 0.3, fusel oil and EtOH 91.4. Fusel oil from fruit is the best, because its value is detd. by the AmOH content. Fusel oil from fruit contains BuOH . A considerable no. of samples analyzed contained abnormal products, such as SO_2 , formaldehyde, fatty acids and amines. Fusel oil from fruit contains an appreciable amt. of compounds boiling higher than AmOH , in 1 case 13.4%. E. M. STOKES

Production of alcohol from spent mohwra (*Bassia latifolia*) flowers. D. L. SARA SRABUDINE AND V. G. PATRARDIAN. *Poona Agr. Coll. Mag.* 22, 45-9 (1930).—The spent mohwra flowers resulting from distillery operations are a nuisance in certain Indian localities. Approx. 11% of total sugar may be obtained from the carbohydrates in the spent flowers by heating the flowers with 0.05 N H_2SO_4 for 4 hrs. under 3 atm. pressure. By diluting the ext. from the original flowers with that from the spent flowers the added sugar ferments completely, giving a proportionately higher total yield of alc. from a given wt. of flowers. Under these conditions neutralization of the H_2SO_4 before fermentation is unnecessary. K. D. JACOB

The composition and determination of the barley proteins. III. Fourth report on barley proteins. The proteins of barley during development and storage and in the mature grain. L. R. HUNTER. *J. Inst. Brewing* 36, 336-40 (1930); cf. *C. A.* 24, 1929.—Analytical data and graphical presentation are given of the proteins of varieties of English barleys, in various stages of maturation, and in mature grain before and after storage. After allowance is made for the various standard errors of the data, it is possible to deduce that the relationship between the quantities of the individual proteins and the total N is characteristic of the variety. Within each var. variety, the weights of the individual proteins are simple regular functions of the total N content per thousand corns. Immature grain, if given sufficient time, reaches the natural equilibrium and the proportion of the various proteins will be the same as in mature grain of the same total N content per thousand corns. It is suggested that in pure grain the protein proportions may be represented in amts. per thousand corns as follows: $S = aN$, $G = bN$, $H = pN + qN^2$ and $X = rN$, where S = salt sol N, H = harden N, G = gluten N and X = total N per thousand corns. The values of the factors (a , b , p , q and r) are regarded as characteristic of the variety. PETER J. F. WEBER

Institute of brewing research scheme. H. LLOYD HIRD. *J. Inst. Brewing* 36, 263-8 (1930).—A report is made by the research organizer covering the autumn and winter months of 1929-1930. PETER J. F. WEBER

Preliminary note on barley proteins. The influence of the method of grinding on the estimation of barley proteins. G. HOFMAN-BANG. *J. Inst. Brewing* 36, 381-8 (1930).—Before making exts. of different proteins in Danish barleys, the reliability of Bishop's method (cf. *C. A.* 22, 3729, 24, 1929) was investigated. The proteins were detd. on barleys ground to varying degrees of fineness as ascertained by shaking in a screen grader. To obtain consistent results and reach a const. value for the various proteins, the barley had to be ground in a ball mill so that 85% passed a sieve with meshes of 0.253 mm. PETER J. F. WEBER

Algerian musts of the 1930 vintage. E. G. FILAUDEAU AND A. BONIS. *Ann. fals.* 24, 5-8 (1931). cf. *C. A.* 24, 2231.—A discussion of the results (not given in detail) found on examn. of 123 representative samples from the Department of Oran, 271 from the Department of Alger and 90 from the Department of Constantine, from which it is concluded that the quality of Algerian wines of the 1930 vintage is, on the whole, satisfactory, and even slightly better than that of the 1929 wines. H. PORTER, SABATIE AND WALLAERT. *Ibid.* 8-11.—A brief discussion of the quality of the 1930 vintage in Algeria, based on the analysis of 253 samples, from which it is concluded that the constitution of the wines should be fairly well balanced, and the alc. content higher than in 1929. A. PAPIREAU COURTES

New apparatus for the rapid determination of the volatile acidity of wine. MARCEL MARTIN. *Ann. fals.* 24, 36-7 (1931).—The simplest method for detg. the volatile acidity of wine consists in subtracting non-volatile from total acidity, but in order to obtain consistent results it is essential that the evapn. of the volatile acids be always carried

out under strictly identical conditions. The app. described consists of an electrically heated water bath, the steam from which is passed through the sample while the latter is being heated on the bath. The construction is such that the quantity of steam passed through the sample is always the same. A. PAPINEAU COUTURE.

Dry, unpressed wine and natural wine. J. WERDER AND C. ZACH. *Wein u. Rebe* 11, 272-6(1929).—Dry, unpressed wine, but not natural wine, after decolorization by animal charcoal, shows a blue luminescence in filtered ultra-violet light. B. C. A.

Rum fermentation. MARTIN FICKER AND STEFAN SZÜCS. *Zentr. Bakt. Parasitenk.*, 2 Abt., 82, 199-211(1930).—The aroma produced during the fermentation of rum is due to the action of acetic, butyric and lactic acid bacteria. JOHN T. MYERS.

Alcohol production for motors in Brazil (FREISE) 21. Polarographic studies with dropping mercury cathode. XVI. Electroreduction of AcH (SMOLEX) 4. The antagonistic substances formed during bacterial fermentation (KLEIN) 11C.

Rectifying alcohol. HEINRICH PROSKER. *Ger.* 515,195, Dec. 21, 1928. Operative details are described. The principal features are the maintenance of an excess pressure of about 0.06 atm. in the still, and the use of a thermometer with a movable scale, which is adjusted in steam before the disto.

Acetyl methyl carbinol. T. H. VERHAAR. *Brit.* 337,025, June 26, 1929. A mash, such as a saccharified potato mash contg. suitable N compds., a phosphate such as superphosphate and a carbonate such as CaCO_3 , is fermented with bacteria such as *Aerobacter aerogenes* which under normal conditions produce 2,3-butylene glycol, and the mash is subjected to aeration insufficiently strong to prevent the formation of butylene glycol and the aeration is continued after the main fermentation period until at least most of the butylene glycol formed has been converted into acetyl methyl carbinol. Volatile fermentation products may be recovered by suitable washing devices, and the acetyl methyl carbinol produced may be converted into diacetyl by the action of oxidizing agents such as FeCl_3 .

Acetic, butyric and propionic acids by fermentation. DISTILLERS CO., LTD., and H. B. HUTCHINSON. *Brit.* 337,153, Sept. 30, 1929. A dild. molasses soln. or other suitable sugar soln. of low concn. (various examples of which are given) is fermented with thermophilic organisms (which may be procured from animal excrement or from decomposed cellulosic material) and stronger sugar soln. is added under regulated conditions such that the concn. remains sufficiently low to prevent the formation of non-volatile fatty acids. Numerous details of procedure are described. Cf. C. A. 24, 2543.

Fermentation. OLGA K. SAUER NÉE BUCHHOLZ. *Ger.* 513,511, Dec. 20, 1927. Yeast and bacteria preps. are rendered stable and capable of germination by mixing to a paste or emulsion with fats at temps. not injurious to the organisms, without subsequent drying. Petrolatum, liquid petrolatum and cacao butter are mentioned as suitable fats.

Promoting fermentation reactions with activated carbon. DARCO SALES CORP. *Brit.* 336,207, April 2, 1929. The yield and speed of fermentation reactions (as in the production of alc., fusel oils and vinegar) are improved by adding to the fermenting mass (which may comprise a sugar or molasses soln. and suitable microorganisms) an active adsorbent agent such as activated "vegetable char" together with org. nitrogenous food substances such as a protein material. The use of "Darco" (produced from lignite) is particularly mentioned. *Brit.* 335,972 relates to similar processes in which the microorganisms and active charcoal are first added (preferably intimately assocd.) to a soln. of medium d. of the material to be fermented and, after the d. has been reduced by the fermentation, it is increased above the initial d. by the addn. of a soln. of high d. of the material to be fermented. An example is given of the fermentation of a molasses soln. by yeast.

Fermenting sacchariferous liquids. EUGEN STICH. U. S. 1,792,450, Feb. 10. In processes such as alc. and yeast production, an app. is employed in which aeration is effected through porous diaphragms of a special degree of porosity (suitably arranged in 2 layers of material, the upper of which, in contact with the liquid, has pores of a diam. of 0.16 to 6 μ while the lower layer has pores of larger diam.). Diaphragms of this particular character are stated to be suitable for producing better yields with a reduced air supply.

Pasteurizing apparatus for beer, etc. OSKAR KARY. *Ger.* 516,458, June 5, 1928.

17—PHARMACEUTICAL CHEMISTRY

W O EMERY

Contribution of the pharmacist to the progress of chemistry and other sciences. GODOY TAVARIS. *Bol assoc brasil pharm* 11, No. 11, 13-22(1930)—Historical review. CARLOS A. ARLEDO

Corrosive sublimate tablets. CARLOS H LIBERALI. *Bol assoc brasil pharm* 11, No. 11, 31-7(1930)—The addn. of NaCl to increase the soly. of $HgCl_2$ as recommended by many pharmacopoeias is not advisable. Complex formation decreases the antiseptic action which is proportional to the concn. of Hg^{++} ions. L. recommends the addn. of tartaric acid which does not interfere with the dissocn. of the mercuric salt. C. A. A.

A new medicine for the treatment of piropilosis. L. TOBAC. *Bull of Congo Belge* 21, 4-8(1930)—A new medicine called "Piropil" has been recommended as a cure for cattle piropilosis. It is a mixt. of trypan blue and Na cholate one of the acid constituents of bile. This medicine may be used up to concns. of 4% and in quantities from 50 to 100 cc. A complete sterilization of the blood is not obtained but this is desirable as the presence of parasites in the blood and resistance to infection. J. R. ADAMS

Beauty creams. FRED WINTER. *Am Perfumer* 24, 232, 234(1929)—A review. E. J. C.

The measurement of the adsorption capacity of medicinal charcoals. HERBIE LAUCKER. *Ann Hochsch* 9, 2296-3000(1930)—Animal charcoals are, in general, better adsorbents than plant charcoals. In comparing the adsorbing power of a large series of charcoals, with methylene blue, morphine and $HgCl_2$ as test adsorbates, a fairly good correlation was obtained, particularly between the last 2. Occasional discrepancies do occur, e. g., a strong adsorption of methylene blue, with weak adsorption of the other 2 by the same charcoal. H. EAGLE

Components of the drug "Salpamist." C. GRIEBEL AND G. STERNOFF. *Arch Pharm* 269, 37-49(1931)—The Last Indian drug "Salpamist" (probably derived from *Allium macleanii*) contains H_2O 9.25, ash 3.75, free acid (as malic) 1.00, fructose 2.04, polysaccharide (calcd. as fructose) 69.63, saponin (in the alc. ext. as calcd. from the difference) 3.97, Et_2O ext. 0.60, insol. 11.00%. Hydrolysis of the saponin (C₄₈H₈₀O₁₆) yields in addn. to sapogenin (C₃₀H₄₈O₈), galactose, arabinose, methylpentose and galacturonic acid. The hemolytic index is 6700, the fish index 13,300. The polysaccharide gives on hydrolysis mainly fructose and is, therefore, to be regarded as fructotosan (C₆H₁₂O₅)_n. W. O. EMERY

Liquor plumbi subacetatis fortis—its preparation and assay. PHILIP A. BERRY. *Chem Eng Mining Rev* 22, 421-3(1930)—Liq. plumbi subacet. fort. owes its virtue to the fact that it is a highly basic Pb soln. No standards requiring a high degree of basicity have been previously suggested, neither has any method of estn. been put forward to ensure this. The following method is proposed for estg. both the Pb and Ac radicals. To a known wt. of sample add a known excess of standard H_2CrO_4 soln., make up to a definite vol., filter and in an aliquot portion of the filtrate det. the excess of oxalic acid by titration with $KMnO_4$. In another aliquot portion det. total acidity by titration with 0.1 N NaOH. The difference gives AcOH. The detn. is not affected by impurities present to an extent exceeding 50% of that permitted by the B. P. To prep. the liquor triturate 320 g. powd. PbO to a smooth paste with 100 ml. distd. H_2O transfer to a bottle, wash in the rest of the PbO with 650 ml. H_2O , shake well, add 240 g. AcOH, shake vigorously for several min. and again at intervals for 24 hrs., filter and pass through the filter water to make up to 1 l. An excess of PbO results in the sepn. of a highly basic Pb acetate from the soln., with a resultant loss of both Pb and Ac radical. The white deposit formed in making this soln. is composed of basic Pb carbonate together with basic acetate of the compn. $Pb(CH_3CO_2)_2 \cdot 2PbO \cdot 4H_2O$. The d. of the official liquor is at present too definite, a range of 1.275-1.300 is suggested. The method of assay proposed is equally suited for Pb acetate. W. O. E.

History of Austrian Pharmacopoeias. OTTO ZEXERT. *Pharm Monatsh* 12, 2-4, 22-5(1931)—An address describing in word and picture certain of the earlier Austrian pharmacopoeial documents and apothecaries. W. O. E.

Arbutin content of certain known Encaceae. L. ZECHNER. *Pharm Monatsh* 12, 4-6(1931), cf C. A. 24, 685.—Some 15 encaceous plants were examd. with respect not only to their arbutin but also their hydroquinone content. The greatest amt. of arbutin occurs in *Arctostaphylos*. W. O. E.

Electrometric studies of complex formation. I. The citrates of bismuth. C.

MURTON *Quart J Pharm Pharmacol* 3, 561-74 (1930).—From the results of an electrometric study in which the glass electrode was employed, it is concluded that: The view that Bi citrate is a complex "bismuthyl citric acid" is incorrect. As ordinarily prep'd, the normal citrate is obtained. In the presence of nickel citrates or bicarbonates, as in the Cowley process, sparingly sol basic complexes are formed, the comp'n of the complex cation being a function of the pH . The process of "neutralization" by alkalis consists in the reversible formation of a series of complex cations, the final product being the neutral, sol and stable 4:1 complex, $3Bi(OH)_3 \cdot 12CaH_2O_2$. "Bismuth and ammonium citrate" is a mixt of this basic salt and of NH_4 citrate in the mol proportions 1:3. The methods given in the B. P. (1914) and the Brit Pharmaceutical Codex (1923) for the prep'n of solns. of "bismuth and ammonium citrate" are unsatisfactory, and suggestions are offered for the prep'n of non precipitable neutral solns. of definite comp'n. W. O. P.

Calcium phosphate B. P. NORR L. ALLPORT AND T. TUSTING COCKING *Quart J Pharm Pharmacol* 3, 575-83 (1930). Analyses of several com. samples of Ca phosphate showing the variable proportions of CaO and P_2O_5 are given. The reaction of $Ca_3(PO_4)_2$ is pH 9.2, and that of $Ca(H_2PO_4)_2$ is pH 4.5. Previous workers have shown that $Ca_3(PO_4)_2$ is slowly decomp'd by H_2O to form $3Ca_2(PO_4)_3 \cdot Ca(OH)_2$. Evidence is here adduced which suggests that, under certain conditions, a salt is formed with a still higher Ca content, and having the approx. comp'n $2Ca_3(PO_4)_2 \cdot Ca(OH)_2$. Expts. have been made which indicate that it is impracticable to manip. $Ca_3(PO_4)_2$ by a process of ppt'n in the presence of H_2O . W. O. P.

Critical observations on the directions for glass testing of the German Pharmacopoeia, 16th edition, L. KROHN *Sprechsaal* 20, 350 (1927), *J Soc Glass Tech.* 13, Abstracts 173-4, cf. C. A. 21, 2047. The German Pharmacopoeia directs that medical glasses be $\frac{1}{2}$ filled with an aq. soln. contg. 1 cc. of 0.1 N HCl per l. and 5 drops of Me. red soln. and heated for $\frac{1}{2}$ hr. in a boiling water bath. After this time the red color of the soln. should not have completely disappeared. For ampoule glasses for solns. of alkaloid salts, 5 g. of the coarsely crushed glass is heated for $\frac{1}{2}$ hr. in a boiling water bath with 100 cc. of water, 0.3 cc. of 0.01 N HCl and 1 drop of Me. red soln. in a Jena glass flask previously rinsed out with boiling distil. water. After this time the red color of the soln. should not have completely disappeared. Tests for pipets, dropping glasses, wide neck glasses and app. glass should be given. Variable results can be obtained by heating the glasses in the steam or in the boiling water, by loosely closing the necks with cotton wool or by securely sealing with cork stoppers, and by having the vessels standing or floating in the water. Preliminary cleansing of the vessels with distil. water at the ordinary temp., to remove dust and grease particles is necessary and ought to be precisely expressed. For pipets, dropping bottles and app. requiring after-working in the blowlamp an intermediate or semi hard glass is suggested. The test for ampoule glass should be more precisely defined, particularly with reference to the size of the crushed glass particles, which should be sieved to a definite uniform size and freed from dust by washing with abs. alc. The narcotine HCl test is superior to any other durability test, because of the extreme sensitiveness of the reagent and the fact that an estn. can be made in the course of an hr. without the necessity of applying heat. G. G.

Rapid volumetric determination of nicotine by means of silicotungstic acid. B. KARANSKY, *Z. an. Chem.* 83, 107-14 (1931).—Nicotine forms a difficultly sol. salt with silicotungstic acid of the formula $2C_{10}H_{14}N_2 \cdot 21H_2O \cdot 12WO_3 \cdot SiO_2 \cdot 5H_2O$ in faintly acidic solns. If, after filtering off the salt, it is decomp'd with caustic alkali and the freed nicotine distil. with steam, the distillate can be titrated with acid. The ppt. can also be weighed. The modification of the method now proposed consists in first distg. off the nicotine with steam and titrating the distillate with silicotungstic acid, stopping at the point where no further clouding is produced by adding the reagent. A 12% soln. of the reagent was used in the expts. described. To carry out the detn. in a sample of tobacco, weigh out 8-20 g. of the material into a 500 cc. round bottomed flask, add 150 cc. of water and 50 g. NaCl. Mix well to a paste and add a suspension of 5 g. MgO in a little water. Distil with steam into a 500 cc. measuring flask contg. a little concd. HCl. After distg. for 1.5-2 hrs. the distillate should show a neg. test for nicotine when tested with silicotungstic acid. Fill the flask up to the mark, mix well and transfer, with a pipet, 100 cc. into a 200 $\frac{1}{2}$ cc. flask. Add 2 g. of solid NH_4Cl and titrate. Standardize the reagent by means of pure nicotine soln. which has been subjected to the same treatment. The analysis of 11 samples of tobacco and nicotine exts. indicates that the method gives satisfactory results. W. T. HALL.

A note on three oil-bearing grasses from Burma. D. RUSBY, *Agr. J. India* 25,

327-9(1930)—A yield of oil amounting to 0.26% of the dry matter was obtained by the steam distn. of fresh *Cymbopogon clandestinus*, a grass which grows abundantly in dry rocky open spaces on hillsides in central Burma. The oil was of a light straw-color and had a pleasant odor somewhat resembling that of ginger grass oil. The oil tends to form an emulsion with water and a redistn. is necessary to effect complete sepn. A new species of annual *Cymbopogon*, smaller than *C. clandestinus* and differing from it in having glabrous pedicelled spikelets, yielded upon steam distn. 0.72% of the dry matter as a pale straw-colored oil, which tended to emulsify in water. Steam distn. of *C. virgatus*, an undescribed perennial species of *Cymbopogon*, yielded 0.55% of an oil having a rather unpleasant odor somewhat resembling that of Botha grass oil. The oil had the following properties: $d_{20}^{20} 0.9566$, $n_D^{20} 1.49107$, and value 1.9, ester value 6.5, ester value after acetylation 72.8, and alc. content calcd. as $C_{21}H_{42}O_2$ 21-22%.

K. D. JACOB

Studies of chemical sterilization and bacteriostasis. I. Standard inhibition control for germicidal tests. J. W. H. HILL. *Am J Pub Health* 21, 192-3(1931)—The following method of making inhibition controls is proposed by which valid germicidal tests may be made by any accepted method. Let *D* represent the no. of cc. of drug employed in the test, *B* the amt. of inoculum employed, *S* the amt. of the mixt. of *D* + *B* used for sub-culture and *T* the no. of cc. of broth used for sub-culture. The inhibition control then consists of the following steps: (1) Mix 1 *D* and an amt. of sterile broth equal to 1 *B*. (2) Transfer the equiv. of 1 *S* of this mixt. to 1 *T*. (3) Mix the equiv. in sterile H_2O of 1 *D* with 1 *B*. (4) Transfer the equiv. of 1 *S* of this mixt. to the 1 *T* previously used in step (2). Such inhibition controls should be made with every diln. of drug employed and with every type of organism used. It is wise to run such controls in triplicate to allow for unavoidable exptl. variation. J. A. K.

Pyrethrum. II. Culture, yield and economic future. J. CHEVALIER. *Bull. us pharmacol* 37, 235-9(1930). III. Industrial and pharmaceutical preparations. The estimation of the activity. *Ibid* 423-31; cf. *C. A* 24, 2804—The pyrethrins I and II are esters of pyrethrolone, a ketonic alc., and of chrysanthemum monocarboxylic acid and chrysanthemum methylcarboxylic acid. The flowers of pyrethrum contain at an av. 2.5-3 g pyrethrin per kg. at a max. 5 g. Stems contain $1/10$ to $1/5$ as much. The industrial exts. are solns. in hydrocarbons or H_2O -sol. emulsions. The pyrethrins can be isolated as semicarbazones, but this method is not suitable for analysis. The extn. with $EtOH$ or $MeOH$ destroys the active substance. Chlorinated org. solvents are not satisfactory. Pyrethrin I is more efficacious than II. In the practical test on insects, different preparations show a large variation of activity. The species of the animal and the temp. have much influence. C. recommends the test on the frog. About 0.2 mg. pyrethrin injected into the lymphatic sac will kill the animal within 3 hrs. The test is unreliable at low temp. The method used for the extn. of male fern by detn. of the effect on fish showed pyrethrin active in dilutions 1:100,000. This is too sensitive for practical purposes. Com. preps. should be tested on the class of animals they are sold for.

A. E. MEYER

The culture of valerian and burdock. G. HINARD AND M. LE PRADRE. *Bull. sci. pharmacol* 37, 479-77(1930)—The compn. and growth of the plants is considered in relation to the compn. of the soils in different parts of France and Belgium. A. E. M.

Essential oil of *Asarum sieboldii* var. *seoulensis* Nakai. L. T. KAKU AND T. KONDO. *J. Pharm. Soc. Japan* 51, 8-17(1931). German abstr. 3-7—K. and K. isolated from *Asarum sieboldii* var. *seoulensis* Nakai an essential oil (yield 2.21%) (I). I has the following properties: $d_{20}^{20} 0.9088$, $n_D^{20} 1.5141$, acid no. 1.85, sapon. no. 28.4, sapon. no. after acetylation 56.1, $\alpha_D^{20} -0.75$ and MeO 21.4%. I contains palmitic acid (II), a substituted phenol (III) with mol. formula of $C_{21}H_{41}(OH)(OC_2H_5)(OC_2H_5)$. III is under further investigation. I, after removal of II and III, was fractionated under reduced pressure. The constituents isolated are *l*-pinene, a ketone b.p. 80-110° (semicarbazone, m. 185°, which, on treating with HCl , gave a free ketone, b.p. 200-20°, $d_{20}^{20} 0.9562$, $n_D^{20} 1.5095$, not purified further, which resembled eucarrone), *safrrole*, which on oxidation gave piperonylic acid, and *methyl Eugenol*, $C_{11}H_{14}O_2$, b.p. 127-9°, $d_{20}^{20} 1.0443$, $n_D^{20} 1.5340$, which on oxidation gave veratric acid, m. 180°.

F. I. NAKAMURA

Common diabetic drug. IV. Constituents of the bark of *Aralia chinensis* L. var. *grabescens* (3). S. KUWATA. *J. Pharm. Soc. Japan* 51, 17-31(1931); German abstr. 7-9, *C. A* 24, 656—From the alc. ext. of the bark of *Aralia chinensis* L. var. *grabescens* K. isolated α -taralin (I) and β -taralin (II). I is sol. in alc., but insol. in H_2O and EtO , it contains a large amt. of ash (chiefly Mg , Ca and Na). It decomp. at 202-4° after electroanalysis. Analysis indicates the mol. formula of $C_{21}H_{32}O_{10}$. Hydrolysis takes

place according to the equation: $C_{42}H_{70}O_{11} + 3H_2O = C_{42}H_{70}O_8 + 2C_{12}H_{22}O_{11} + C_{18}H_{32}O_{16}$ (*Allyl- α -fructan* decomps. 143°). Catalytic reduction of I gave glucose (decomps. 220°) which shows strong acid and Liebermann's reaction. The hemolytic action of I is similar to that of saponin of Kahlbaum and senegin of Merck. The action is weakest in a neutral soln. P. I. NAKAMURA

Digitalin of Nativelle and digitoxin. V. HASENFRATZ. *Compt rend* 192, 366-8 (1931).—II. attempts to prove that the digitalin of Nativelle and the digitoxin of Schmeideberg are the same definite chem. substances by expts. showing that these comds. have identical products of hydrolysis. The reaction of Windaus and Stein (C. A. 23, 1646) representing the hydrolysis of digitoxin into digitoxose and digitoxigenin ($C_{42}H_{70}O_{11} + 3H_2O \rightarrow C_{42}H_{70}O_8 + 3C_{12}H_{22}O_{11}$) also represents, according to II, the hydrolysis of digitalin of N. It is pointed out that digitalin of N. and digitoxin have the same toxicity (cf. Raymond Hamet, C. A. 23, 2757). ALBERT L. RAWLINS

Experimental studies on heart tonics. IV. The main factors of digitalis standardization with a new assay method. WILLIAM NYRI and LOUIS DUBOIS. *J. Pharmacol.* 40, 373-401 (1930) (C. A. 24, 5381)—Rabbits are anesthetized by intravenous injection of barbital derivatives, and the drug to be tested injected into the jugular vein. Blood pressure tracings are made (in some instances using ouabain to supplement the prepn.) until the pressure drops to zero. The prepn. tested may subsequently be used without the further complication of inaccuracy by diln. C. RIGGALL

Trypsin preparations suitable for the prevention of adhesions. ROBERT P. WALTON. *J. Pharmacol.* 40, 403-11 (1930)—Bacteria-free enzyme prepns. could be obtained by pressure filtration of water-glycerol solns. of trypsin, and from this, by pptn. with alc. and ether, a dry sterile powder. These prepns. were found to be effective in preventing peritoneal adhesions in dogs. C. RIGGALL

Analysis of disinfectants. LESLIE HART. *Susp* 5, 111-5 (1930)—The methods of the Food and Drug Administration St. Louis station are given for the following detns.: water, alkalies, phenols, neutral oils and fatty or resin acids, mineral oils, mineral oil in creosote oil, Liquor Cresolis Compositus, unspnd. material and fatty acids, total alkali, glycerol, saponin, cresols, water in pine oil emulsions, pine oil and fatty and resin acids. E. SCHREIBER

Hypophosphites. I. H. ALCOCK. *Chemist and Druggist* 113, 613-4 (1930).—Detailed criticisms are given on the monographs of the various hypophosphites of the new Brit. Pharm. S. WALDROTT

Glycerinum acidi borici, Brit. Pharm. J. F. BOYLE. *Pharm. J.* 125, 453 (1930).—The present Brit. Pharm. requires a soln. of 300 g. B(OH)₃ in 450 g. glycerol to be evapd. to 500 g. at a max. temp. of 150°. On a large scale a higher temp. is needed for evapn., which causes darkening of the product. A simplification of the proportions instead of an increase as proposed for the new Pharm. is suggested. S. WALDROTT

Unguentum hydrargyri compositum. J. H. RIDGWAY. *Pharm. J.* 125, 453 (1930)—The formula ung. hydrarg. Brit. Pharm. 400, cera alba, Brit. Pharm. 50, camphora 100, adeps benzoatus 450 is recommended in the place of the Brit. Pharm. formula in which the yellow beeswax imparts to the prepn. an unsightly greenish color. S. WALDROTT

Comments and criticisms, British Pharmacopoeia revision. ANON. *Pharm. J.* 125, 453, 483, 511, 536, 561, 647-8, *Chemist and Druggist* 113, 553-4, 587-8, 613-4, 614-5, 635-6, 746, 770-1 (1930), cf. C. A. 25, 772.—More or less detailed comments were made thus far by 31 pharmacists. S. WALDROTT

Acidum carboicum liquesfactum. Infusions. B. G. W. HOARE. *Pharm. J.* 125, 483 (1930).—The acidum carboicum liquesfactum proposed for the new Brit. Pharm. is said to contain 30 grains in a fluid drachm. A sample made according to the new formula had a sp. gr. of 1.042, hence it contains a little below 46.5 grains per fluid drachm. Fresh infusions are superior in aroma and in efficacy in certain cases, to comcd. infusions, to introduce these would be a backward step. Infusions of ergot, broom, chiretta and rhatany might be deleted. S. WALDROTT

Syrupus ferri phosphatis compositus. F. GOLDBY. *Pharm. J.* 125, 561 (1930).—A decoction of cudbear instead of cochineal is recommended as a permanent, bright coloring agent. Possibly a sol. powd. ext. of cudbear may be used. S. WALDROTT

Note on linimentum camphorum. J. F. DUCKWORTH. *Pharm. J.* 125, 647 (1930).—Peanut and ordinary cottonseed oil are not suitable for the prepn. of a satisfactory camphor liniment since the latter prepd. with these oils become solid or opaque at 0°. The non-freezing cottonseed oil is suitable and should be adopted in the new Brit. Pharm. S. WALDROTT

Liquor carminal, Brit. Pharm. Codes. J. F. BOYLE AND J. H. FRANKLIN. *Pharm.*

J 125, 648(1930) — Evapn. of the NH_4OH soln. of carmine to near neutrality, as directed by the Codex, causes much pptn. of the coloring matter. It is suggested that after evapn. until the NH_3 is expelled, just enough soln. of NH_4OH be added to redissolve all of the pptd. carmine, or dissolve the carmine in NH_4OH and preserve the soln. with glycerol
S. WALDBOTT

A qualitative comparison of various digitalis substances. HARVEY GOLD, WM. HITZIG, BEN GELFAND and HERMAN GLASMAN *Am Heart J.* 6, 237-54(1930) — The substances used were tincture digitalis of an American grown leaf, the same of a German grown leaf, the chloroform fraction of digitalis, tincture of adonis, ouabain and digitoxin. All have the same qual. cardiac actions, but there are extraordinary individual differences in response to the same prepn.
R. C. WILLSON

Standardization of the trypanocidal activity of some aryl derivatives of arsenic acid (LAUNOV, ENGLER) 1111. Chloroformates of alkaloids (TCHAKIRIAN) 6. Explosions occurring during the use of C_4H_8 (JOHNSON, CABOT) 24. Gases from crude oil cracked in the vapor phase [for use in preparing perfumes] (MARKOVICH, PIGULEVSKII) 22.

CHIRIS, A. Contribution à l'étude des huiles essentielles. Grasse. Antomar Chir. 55 pp. Reviewed in *Chimie & Industrie* 24, 1507(1930)

Synthetic drugs. I G FARBERND A-G Brit 337,136, Oct 8, 1928. Acetates of β -aminoacridine or its substitution products, which are readily water sol. and suitable for therapeutic use, are obtained by treating the base with HOAc and isolating the product. Examples are given of the prepn. of acetates of 7-ethoxy-3,8-diaminoacridine, 9-aminoacridine, 7-methyl 9-aminoacridine and 3,8-diaminoacridine. 7-Methyl 9-aminoacridine is prepd. from β -chlorobenzoic acid and p -toluidine by a process similar to that described in Brit 170,038

Synthetic drugs. I G FARBERND A-G (Hubert Oesterlin, Karl Streitwoll and Alfred Fehle, inventors) Ger 517,002, Apr 8, 1929. α -phenolglycinamide-arsonic acids and their substitution products are prepd. from α -aminophenol-arsonic acids and their substitution products by the customary methods. Thus, 3-amino-4-hydroxybenzenearsonic acid may be treated with CH_3O and KCN to yield the nitrile of 4-hydroxybenzene-1-arsonic-3-aminoacetic acid, from which the acid and its amide are obtained in turn. Other examples are given also.

Therapeutic preparations. SALO BRAGA, Ger 513,694, April 24, 1929. Antisymphylis spirochetes prepns. are obtained by the intrapleural, intraperitoneal or intravenous injection of animals with lipid or lipid albuminous substances, nucleoproteins, or albumin, inoculating them with weakened and then with fully virulent syphilis spirochetes in the breast or abdomen and ear vein, and working up the exudations from the breast and abdomen with the ext. from the lymphatic glands, milk, ovaries, etc., and with the blood serum.

Medicinal aqueous-oil emulsions. CLAUS N STRAUCH U. S. 1,791,878, Feb 10. An emulsion suitable for injection into the tissues is prepd. by dissolving a cholesterol (such as metacholesterol) and beeswax in olive oil, and adding a water sol. drug such as insulin. Myrsein may be added as a stabilizing ingredient.

Complex iron compounds of aliphatic hydroxycarboxylic acids. I G FARBERND A-G Brit 335,965, June 5, 1929. Acids such as hexantritollic acid, gluconic acid, trihydroxyglutanic acid, glucuronic acid, saccharic and mucic acids and their isomers or salts are treated with Fe or an Fe compd. and the reaction mixt. is simultaneously or subsequently neutralized with a basic substance such as alkalis or NH_3 or org. bases. Water-sol. therapeutic compds. are obtained, and details are given of the production of Na ferrous and ferric gluconates, K-Na ferrous and ferric saccharates, quinizarin K ferrous saccharate and Na Fe mucate.

Anesthetics. I G FARBERND A-G (Max Bockmühl and Leonhard Strin, inventors) Ger 516,285, Jan 16, 1929. Alkamine esters of halophenyl carbonic acids are prepd. by condensing halides of halophenyl carbonic acids, or their substitution products free from N, with aliphatic amines having a tertiary N atom. Thus, p -chlorophenyl carbonic acid chloride (prepd. from p -chlorophenol and COCl_2) may be caused to react with $\text{Et}_3\text{NC}_2\text{H}_5\text{OH}$, yielding $\text{ClC}_6\text{H}_4\text{OC(O)OCH}_2\text{CH}_2\text{NEt}_3\text{HCl}$, m. 132° . Other examples are given also.

Extracting organic substances. I G FARBERND A-G Fr 694,602, April 23, 1930. Org. substances are extd. from animal or vegetable materials, e.g., alkaloids from roasted coffee or tobacco, or sugar from beet roots, by means of liquid NH_3 .

Hormones. SCHERING-KAHLBAUM A-G Brit 336,470, Dec 6, 1928 For purification of crude hormone preps such as those derived from animal liquids or organs, yeast or maize shoots, impurities of an ester like character are saponified by Twitchell reagent, Darting saponifier, lipase of castor seed, pancreas lipase or other catalytically acting agent. The saponified material may be extd with MeOH, the ext. neutralized with alc. $\text{Ba}(\text{OH})_2$, the filtrate concd and then taken up in water with ether, and the hormone obtained by evapg the ethereal soln.

Hormones. SCHERING-KAHLBAUM A-G Brit 336,471, Dec 1, 1928 Liquids contg hormones such as urine of pregnant mammals or an aq ext. of the anterior lobe of the hypophysis are mixed with a water sol. org. liquid such as alc. in quantity insufficient to ppt the hormone but sufficient to ppt impurities. The hormone is subsequently pptd as by addn of ether, CHCl_3 or ethylene chlorohydrin.

Tuberculosis vaccine. CARL OTTO Ger 516,403, Dec 7 1926 Pus from a dead tubercular cow is injected into a healthy cow and the latter, at the end of the incubation period, is inoculated subcutaneously with cream cal. The pus produced at the point of inoculation is removed coagulated, and filtered. The residue is mixed with ground periosteal substance and the mixt finally sludged with the filtered pus serum.

Alkaloids. SOC A S FARMACON Tr 694055, April 16 1929 Opium is treated with a solvent for fatty materials, preferably petroleum ether then the soln is agitated with acidulated water, preferably contg HCl . The acid soln is neutralized to ppt. the alkaloids.

Digitalis glucosides. WELLCOME FOUNDATION LTD and S SMITH Brit 337,091, Aug 8, 1929 A new glucoside is obtained by stirring the total glucosides of the leaf of *digitalis lanata* with acetone or methyl ethyl ketone, filtering fractionally pptg the filtrate by adding water, removing the sepd. solid, adding salt to the clear liquid, drying the ppt. thus formed, and extg it with MeOH or EtOH on adding water to the soln. and standing, crystals of the new glucoside are obtained. Various details of purification and auxiliary treatment are given.

18—ACIDS, ALKALIES, SALTS AND SUNDRIES

E. M. STOKES

Purification of caustic soda obtained by the lime process. ALFONS KRAUSE AND WACLAW KILKA *Przemysl Chem.* 15, 6-12(1931)—Caustic soda obtained by the CaO process contains Na_2CO_3 , NaCl , Na_2SO_4 , Na_2SiO_3 and traces of Al_2O_3 and Fe_2O_3 as impurities derived from the raw materials. Their removal by the chem. method is impossible, but it can be accomplished by concn and cooling. Studies on the solubilities of the above salts in NaOH solns. of varying concn. and at various temps. show that the concn. of the NaOH should be increased up to 72-5% at least, higher concn. not to exceed 100% being preferable. Cooling decreases the soly. of all the impurities. For the removal of Na_2SO_4 to a very low figure cooling of the NaOH to about 50° is necessary. The degree of cooling required for the removal of other impurities depends on the concn. It is well to concn. the NaOH in the presence of a definite amt. of CaCO_3 , which forms insol. compds. with Na_2CO_3 and Na_2SO_4 and thus causes their removal from the soln. Among the impurities in the purified soda Na_2CO_3 takes first place and NaCl second. There is relatively little Na_2SO_4 and only traces of Al_2O_3 and Fe_2O_3 . A. C. ZACHARY

Evaporation of electrolytic caustic soda. J A LEE *Chem Met Eng* 37, 494-8 (1930)—An account is given of the continuous evapn. and salt recovery system at the South Charleston plant of the Westvaco Chlorine Products, Inc. The original soln. contains about 8% NaOH and 17.5% NaCl , and the final products from the steam evapn. plant are 53 short tons of NaOH per 24 hrs. in soln. (d. 1.53) contg also about 0.9% NaCl in soln. and none in suspension, and NaCl brine for return to the electrolytic cells. The water evapd. is about 1 million lb. per day, and the steam used, at 10 lb./sq. in. gage, about $\frac{3}{4}$ million lb. Evapn. is in 3 stages, a double effect followed by a single; then follows a continuous crystallizer and Laughlin centrifugal filter, which delivers the final caustic liquor. Swenson evaporators, Dorr thickeners and rake classifiers are employed. The materials of construction used are Ni cast Fe, Ni tubes and Monel metal. B. C. A.

Causticizing soda by Lewig's method and the use of various iron ores for this purpose. V S. YATLOV. *J. Chem Ind (Russia)* 6, 1029-7(1923)—The reaction between

soda and several Fe ores was found to be very sensitive to such factors as time, temp., subdivision of the materials, and make-up of the mix. A change in any one of them requires great readjustment in the others to bring about the same degree of reaction or the same velocity. The surface area of the reacting substances is the most important factor of all. Under the usual factory conditions the currents of hot air often carry away the finest particles of the ore, and heat sometimes produces a fusion of the reactants with resulting diminution of the surface area. In studying the factory phase of this reaction it appears that further work should be concerned with the use of more finely subdivided ore and the use of lower temps. Special lab analyses were worked out to obtain more accurate results than those of routine analysis. A. C. ZACHARY

Causticizing ammonium chloride with magnesium oxide. KOJI NISHIDA AND SOEMAN NAKAMURA. Abstracts from Rept. Central Lab. S. Manchuria Ry. Co 1029, 37-9.—The NH_4Cl soln. must be dil. to 2.34 N for perfect recovery of NH_3 . This soln. uses 3 times as much fuel for distn. V. F. HARRINGTON

Dicalcium phosphate or "precipitate," and its commercial production. REP. WEBERS. *Mech. Horse* 20, 2059-60, 2555-c (1930)—A review of methods of manuf. of CaHPO_4 from superphosphate and bones. E. M. SYMMES

Preparation of zinc oxide by the wet method. TATSUO IRIYE. *Repts. Imp. Ind. Research Inst., Osaka, Japan* 11, No. 16, 56 pp. (1930)—The powd. Zn ore is roasted and treated with H_2SO_4 . The ZnSO_4 thus obtained is treated with NaOH soln. under pressure. The crude Na_2ZnO_2 is freed from other metallic impurities and treated with CO_2 . Pure Zn(OH)_2 is oxidized to ZnO by heating. F. I. NAKAMURA

Catalytic reduction of carbon monoxide under ordinary pressure. VII. Effect of alkalis on iron-copper catalysts. S. KODAMA AND K. FUJIMURA. *J. Soc. Chem. Ind., Japan* 34, Suppl. binding 14-6 (1931) (in German), cf. C. A. 25, 1039.—With a Cu-Fe catalyst prepd. from the nitrates, 30 cc. of oil contg. a small quantity of a substance like vaseline was obtained at 270° from 141 l. of water gas. Addn. of 0.003-0.03 mol. of Na_2CO_3 to the above catalyst increased its initial activity and favored formation of higher hydrocarbons. With catalysts contg. alkalis there is a high gas contraction at first, the activity diminishing gradually. The catalyst may become coated with the higher hydrocarbon. No diminution of activity occurs with alkali-free catalysts. The best catalyst contains Fe 1 atom, Cu 1 atom, Na_2CO_3 0.003 mol. E. M. S.

Solid carbon dioxide from Mexico. J. W. MARTIN. *Ind. Eng. Chem.* 23, 256-8 (1931)—Oil well gas, contg. 95% CO_2 and 5% combustible gas, issues from oil wells in the Quebrache field, Mexico, at 100°F and a pressure of about 1000 lb per sq. in. After purification, the gas is cooled to the liquid state and then expanded to form solid CO_2 , which is shipped by rail to the Gulf Coast and by boat to New York. The plant cost is $\frac{1}{2}$, and the transportation expense, including duty, $\frac{1}{10}$ of the total cost. F. D. ROSSINI

Carbon dioxide in its new field of usefulness. J. C. GOOSMANN. *Ice and Refrigeration* 80, 45-9 (1931)—A review of the origin, liquefaction and solidification of CO_2 . Various methods now in use for the solidification of CO_2 , together with the thermodynamics of each and a tabulation of the power required per ton of solid CO_2 manuf., were described previously (August, 1930). The chronology of CO_2 discovery is given, with a brief history of patent literature. Early methods of producing and compressing CO_2 snow to "dry ice" are described, and the development of the present methods of production is discussed. ALICE W. EPPERSON

Regaining iodine from phosphates. ILANS JOANN. *Chem.-Zig* 55, 86 (1931)—From the analysis of the crude I, it is concluded that this by product is obtained by the Dansk Svovlsyre und Superphosphatfabrik of Copenhagen electrolytically, and the procedure is condemned as non lucrative. J. is of the opinion that by treating the gases evolved from superphosphate manuf. with CS_2 , active C, H_2O_2 and O_2 , the I₂ can be brought by filtration to a sufficient state of concn. to render further purification profitable economically. B. S. LEVINE

Evaporation of natural salt brines. V. P. LUTSEN AND V. M. FURERO. *J. Chem. Ind. (Russia)* 6, 1163-8 (1929)—The phase diagrams of Na and Mg chlorides and sulfates were studied in relation to sea water and lake brines. The processes of evapn. of natural brines with crystn. of NaCl were classified according to the types of Glauber salt, astrachanite, epsomite, hexahydrate and biphasate. A new type of evapn. was established, based on Caspian Sea brine, accompanied by a labile system. NaCl deposits in the astrachanite sector, as a result of which NaCl deposits out in a pure state to the amt. of 30-40%. The labile equil. in the epsomite sector is disturbed by fluctuations of temp., which cause a deposition of epsomite (on the av. found for stable equil.) depending upon areas of crystn. as worked out in the lab. The amt. of

NaCl depends on the relative NaCl content of the dry residue, and to an even greater degree on the coeff. of metamorphosis. The yield of salt declines sharply with increase in the coeff. of metamorphosis. On the basis of industrial observations and data in literature concns. were established at which crystn. of NaCl begins in the various reservoirs under normal conditions and at an av temp of 25° A C ZACHLIN

Salt and its production in history HANNA FREYDANK. *Kali* 23, 145-51, 161-8, 177-81 (1929) E. J. C.

The utilization of bittern. I. The recovery of Glanher's salt. TORARO KATO *Abstract from Rept. Central Lab. S. Manchuria Ry. Co.* 1929, 40-1—One hundred parts bittern and 40 parts NaCl are dild. with water until d = 24° B ϕ . Cooling to -10° gives 4.1% of the Na₂SO₄ in the bittern and cooling to -19°, 70% V F II

Quinquennial review of the mineral production of India for the years 1924 to 1928. Salt. W A K. CHRISTIE *Records Geol. Survey India* 64, 276-87 (1930)—About 1/4 of Indian salt production is from sea water, 1/4 from rock salt and the rest from areas of internal drainage. The salt works are described BORAX. E H PASCOE *Ibid.* 345-7—Borax is not known to occur in India, although the country exports borax obtained from hot spring deposits in Kashmir and salt lakes in Tibet. Sodium compounds (other than salt) W A K. CHRISTIE. *Ibid.* 432-6—Na₂CO₃ (usually) and Na₂SO₄ deposits (occasionally) occur. Soda is evapd. from lake waters A II E

The dehydrating action of coal ash. MUNEYARI TANAKA *Abstracts from Rept. Central Lab. S. Manchuria Ry. Co.* 1929, 36-7—Fuchun coal ash and Japanese acid clay at 200° have action only slightly inferior to AlCl₃ in the formation of anthraquinone derivs. from phthalic anhydride and phenols V. F. HARRINGTON

Driers in soda factories. V G BAKHROV *J. Chem. Ind. (Russia)* 6, 694-8 (1929)—One of the main points of weakness of the usual soda process is the low efficiency of the driers which show a heat utilization for evapn. purposes of only 45-50%. The performance of the drier can be improved by replacing by a centrifuge the continuous filter which feeds the NaHCO₃ into the drier, and shortening the rotary drier itself. A complete heat and material balance is calcd. for an av. plant A C ZACHLIN

Performance of the Kessler apparatus. V N SCHULTZ, U M. VOLCHENKO, I V INFLOVSKII AND L. V. SEPONTZKII *J. Chem. Ind. (Russia)* 6, 923-40 (1929)—A careful study was made of material and heat balance in the Kessler app. in the Degoonin H₂SO₄ chamber process plant. The exceptionally high heat efficiency of this particular app. depends in part on the high temp. of gases entering the saturator. Utilization of heat in the Kessler app. was distributed as follows: concn. of H₂SO₄, 77.9%; outgoing gases 13.2%; loss to surrounding medium 8.9% A C ZACHLIN

Revivification of foam solutions for fire protection with carbon dioxide. C B WHITE. *Chem. & Met. Eng.* 37, 379-80 (1930)—By a few simple adds. to the fire-protection pump the soln. of NaHCO₃ is circulated from the bottom of the storage tank, through a small tower, and back to the upper part of the tank, it is recarbonated in the tower by gas from a cylinder. The method is preferable to adding addnl. NaHCO₃, because if that is done, the Al₂(SO₄)₃ or other acid soln. which must also be added has to be increased *pro rata*, and eventually both will have to be scrapped. B. C. A.

Potash salt recovery in cement burning (Krusk) 20. Apparatus for NH₃ synthesis and like catalytic gas reactions (Ger. pat. 516,249) 1. Sulfonic acids of the fatty series (Fr. pat. 694,692) 10. Manufacture of liquid household insecticides (Weed) 15.

KREZIL, FRANZ. Untersuchung und Bewertung technischer Adsorptionsstoffe. Leipzig. Akad. Verlag. 504 pp. M. 28, bound, M. 30

Hydrocyanic acid. IMPERIAL CHEMICAL INDUSTRIES, LTD., T S WHEELER, II. A. T. MILLS, J. MCAULAY and W B FLETCHER. *Brit.* 335,947, March 27, 1929. In prep. HCN by reaction between hydrocarbon gas and NH₃, as described in *Brit.* 335,585 (C. A. 25, 1642), the reaction mixt. may contain one mol. proportion of NH₃ to one atomic proportion of C, or less, if the mixt. also contains H.

Hydrocyanic acid stabilization. HANS LEHRECKE (to Deutsche Gold & Silber-Scheideanstalt vorm. Roessler). U. S. 1,792,103, Feb. 10. As a stabilizing agent there is added to liquid HCN a small proportion of an ester of a halogen substituted org. acid such as Et bromoacetate.

Nitric acid. E. I. DU PONT DE NEMOURS & CO. *Brit.* 336,638, April 17, 1929. Absorption in water or HNO₃ soln. of gases contg. N oxides (such as those produced by the arc process or by the catalytic oxidation of NH₃) is effected under pressures

of 50-150 lb. per sq. ft. in a specially constructed tower (details of which are described) comprising 10 to 20 sections for gas and liquid contact. The gases are alternately scrubbed with liquid in each section and further oxidized on their passage to the next higher section. Cf. C 1 24, 243.

Nitric acid production by catalytic oxidation of ammonia. E. I. du Pont de Nemours & Co. Pat. 2,622,223, March 27, 1953. In a process involving oxidation of NH_3 to nitric oxide, further oxidation of the nitric oxide to NO_2 , and absorption in water condensed from the oxidized mist, the oxidation and absorption are effected under pressure (suitable 4-10 atm. or higher) and the oxidized mist is cooled quickly to effect separation of water before oxidation of the nitric oxide takes place. An arrangement of app. and various details of procedure are described.

Recovery of nitric acid. MARION G. MILLER (to the Hercules Powder Co.). Can. 328,571, Feb. 24, 1951. Waste gases from nitrating operations are treated for the recovery of HNO_3 by mechanically removing H_2SO_4 mist, introducing NH_3 , and separating substantially pure NH_4NO_3 formed by combination of the NH_3 with HNO_3 vapor of the waste gases.

Sulfuric acid. HENRI N. M. CORSTADT. Fr. 624,080, April 17, 1930. The residual gases from the Pb chambers are practically perfectly dehydrated, and the loss in S is suppressed by substituting the gases to the action of H_2SO_4 only after they have reached a degree of oxidation corresponding to the presence of NO and NO_2 in equimol. proportions, the treatment with H_2SO_4 being then reduced to about 5-10 sec. The gases are given a final washing with alkali carbonate or bicarbonate.

Sulfuric acid. I. G. FARBENFAB. A.-G. Fr. 624,081, May 1, 1930. Contaminants of HNO_3 are eliminated from impure H_2SO_4 containing more than 60% of H_2SO_4 by adding 50% dissolved in an appropriate solvent to the H_2SO_4 , to be purified and treating the mist with a current of gas at a high temp.

Cooler for sulfuric acid. A.-G. GEORG ECKHART'S SALZWERKE UND CHEM. FAB. Ger. 516,278, July 23, 1929.

Residue of sulfuric acids from oil refining. STANISLAW PILAY and JAKOB L. STRADA. Fr. 624,227, Apr. 10, 1930. The residue of sulfuric acids obtained in the refining of oils with H_2SO_4 is neutralized with alk. earth hydroxides and treated with an alkali carbonate to obtain alkali sulfates, which may be treated with mineral acids to obtain the free acids. The products are used as antisept., emulsifying and foaming agents.

Oxidation of ammonia. I. G. FARBENFAB. A.-G. Fr. 624,729, May 1, 1930. NH_3 is oxidized with catalysts of the Pt group, covered with Rh alone or with a mist of Rh with other metals of the Pt group. The Rh is heated to redness to make it adhere better before use as a catalyst.

Apparatus for the catalytic oxidation of ammonia. SOC. ANON. D'ETRAITS-MARITIME. Ger. 516,444, April 17, 1929.

Mono- or di-alkali cyanamides. JEAN MARCOLE. Fr. 624,032, April 17, 1930. CaCN_2 is caused to react with an alkali phosphate in dil. or concd. soln. and at any temp.

Alkali phosphates. METALLGES. A.-G. Fr. 624,677, April 22, 1930. Alkali phosphates, particularly Na_2HPO_4 , are prepd. by heating NaCl or other alkali chloride in a cyclic operation with FeS in the presence of O and steam, the Na_2SO_4 formed is heated with Fe-P to obtain Na_2HPO_4 , and FeS which is used again.

Alkali phosphates. METALLGES. A.-G. (Friedr. Conway von Gusewald, Hans Weidmann and Gerhard Roemer, inventors). Ger. 516,762, June 1, 1929. See Brit. 235,492 (C 1 25, 1643).

Pyrophosphates. METALLGES. A.-G. Fr. 623,913, April 15, 1930. Acid di-Na pyrophosphate ($\text{Na}_2\text{H}_2\text{P}_2\text{O}_7$) is made by evap. under vacuum solns. of phosphoric acid contg. the corresponding amt. of NaCl . Temps. which would cause the formation of mol. metaphosphates are avoided.

Phosphorus compounds. VICTOR CHEMICAL WORKS. Fr. 624,278, April 22, 1930. An app. is described for the manufact. of P compounds, consisting of a combustion chamber and means for leading in gases contg. P.

Anhydrous chlorides. I. G. FARBENFAB. A.-G. (Max Jager, Robert Suchy and Wilhelm Moschel, inventors). Ger. 513,529, Mar. 30, 1929. Addn. to 502,645. Anhyd. chlorides of Li, Ca, Zn and Ce are obtained by the method used in 502,645 (C 1 25, 544) for anhyd. MgCl_2 , i. e., by heating a mist of metal chlorides and oxides of Li, Ca, Zn or Ce. In one example, a mist of CaCl_2 and CaO are heated electrically in a current of Cl_2 and CO or HCl. In another, NaCl and KCl are heated with CeO_2 .

Alumina. MAX BUCHNER U S 1,792,410, Feb 10 Alumiferous material such as clay is treated with HNO_3 insufficient to combine with all the alumina in the material, and the Al nitrate soln is sep'd and decompd by heat Cf C. A. 24, 5947.

Alumina. WM GERTLER Ger 516,278, June 11, 1929 In the manuf. of Al_2O_3 by heating Al silicates with CaCl_2 or like chloride at 650-1000°, the HCl evolved is caused to act again on the reaction mixt when the latter has cooled to below 400°. Ca silicate is thus decompd and CaCl_2 is recovered The reaction product is worked up by extg CaCl_2 with water, and extg Al(OH)_3 from the residue with HCl , the AlCl_3 soln being then worked up in known manner Details are given

Ammonium carbonate. I G FARNEND A-G Brit 336,206, July 1, 1929 Reaction under pressure is effected with excess of liquid or gaseous NH_3 and NH_4HCO_3 , $(\text{NH}_4)_2\text{CO}_3$ or bicarbonates of alkali metals or of Mg (in which latter cases carbonates of the alkali metals or Mg are also formed) Mixts of salts produced may be sep'd by treatment with water, aq or liquid NH_3 , MeOH , etc., and the reactions may be carried out in the presence of water

Diammonium phosphate. I G FARNEND A-G Brit 336,008, July 10, 1929 Di NH_4 phosphate soln is mixed with the dry salt to obtain a moist product which can be dried at a temp. of about 60-80° without loss of NH_3 The soln used is preferably made from H_3PO_4 , still contg some of the H_2SO_4 used to decompose the crude phosphates, evapd to a sp gr of 1.4-1.5 and cooled The contained Ca is thus pptd as sulfate and is removed, the soln being then neutralized The pptd. Fe and Al may be left in the soln

Calcium carbonate. LOUIS PERIN Fr 694,507, Aug 1, 1929 Chemically pure CaCO_3 is obtained by converting the natural carbonate into chloride and treating the chloride with $(\text{NH}_4)_2\text{CO}_3$ which ppts pure CaCO_3 The $(\text{NH}_4)_2\text{CO}_3$ is recovered by treating the residual soln. with CaO and carbonating with CO_2 , the CaCl_2 being then treated as indicated.

Dicalcium phosphate and potassium nitrate. A HOLZ and T VAN D BERDELL Brit 336,692, July 30, 1929 Material contg tri Ca phosphate is treated with a mixt of H_2SO_4 and HNO_3 (preferably in the ratio of 1 to 4 mols.), CaSO_4 is removed from the soln of H_2PO_4 and $\text{Ca(NO}_3)_2$ formed, and the latter is then treated with KOH or K_2CO_3 Di-K phosphate formed is immediately converted into di Ca phosphate, and the latter is removed and dried, the remaining KNO_3 soln may be evapd to dryness, the salt may be mixed with di Ca phosphate contg some CaSO_4 , and the latter may be converted into $(\text{NH}_4)_2\text{SO}_4$ Cf C A 25, 1324

Carbon monoxide. HENRY DREYFUS Fr 694,223, April 19, 1930 Mixts of CO_2 and H_2 , particularly those resulting from the washing waters of industrial mixts. of these gases, are passed over chromite heated to 800-900°, and the resulting gases are passed over hot C (coke, coal, wood, etc.) in any manner used for making CO from CO_2

Carbon dioxide. I G FARNEND A-G Fr 694,871, May 2, 1930 Liquid CO_2 is prep'd by the catalytic disocn under pressure of CO or gases contg it, and condensing under pressure the CO_2 formed Examples are given using Ni as a catalyst

Solid carbon dioxide. KARL FEHRMANN Ger 513,514, Nov 4, 1928 A method and app. are described for regulating the evapn. of solid CO_2 by using the gas pressure developed

Solid carbon dioxide. GFS. FFR LINDE'S FISMASCHINEN A-G Ger 513,528, Apr. 17, 1929 Addn to 463,125 Solid CO_2 is obtained by sep'n of its vapor from a solidified mixt. and resolidification App. is described

Chromium oxide. I G FARNEND A-G Brit 336,671, July 19, 1929 A product suitable for use as a pig. ext. polishing agent or for metallurgical purposes is obtained by igniting (suitably at about 900°) Cr hydroxide obtained by reducing an alkali chromate or dichromate (free from coloring impurities such as Fe) with S in soln. or partial soln. and suspension Alkalies and polysulfides, etc., may be used as initial materials for the reduction, and the color of the ignited product varies from yellow to blue as the temp. of reduction is increased (reduction temps. of 105-120° being suitable) $\text{Na}_2\text{S}_2\text{O}_8$ may be crystd. from the spent liquor and the latter then reused, or part of the S may be recovered by treatment of the liquor with H_2SO_4

Phosphorus oxychloride. J S DUNN, I BRITTS and IMPERIAL CHEMICAL INDUSTRIES, LTD Brit 336,065, Sept. 6, 1929 Ca phosphates such as rock phosphate or double superphosphate are reacted on by phosgene under such conditions as constantly to expose fresh surfaces to the reacting gases (as by agitation with admixed pebbles in a rotary kiln), and the CaCl_2 (formed from CaCO_3 in the raw material) is prevented from forming a protective covering The mass may be washed with water at inter-

vals and reused after drying, and inert solids such as coke or washed residues from a previous operation may be added. Alternatively, the phosphate may be mixed with C activated by treatment with Cl at 1000°, and Cl alone or mixed with CO may be used instead of phosgene. The temp may be raised during the reaction from about 350° to 600° (or higher if Cl alone is used).

Phosphorus oxychloride. J. S. DUNN, F. BAIRAS and IMPERIAL CHEMICAL INDUSTRIES, LTD. Brit. 337,123, Sept. 6, 1929. POCl_3 is produced by reaction of COCl_2 or its equivs. such as CO and Cl, upon natural Ca phosphate such as phosphate rock, which has been subjected to a preliminary treatment to remove other reactive Ca compds. such as CaCO_3 (suitably by conversion of the latter into sol. compds. and washing out these). In a modified procedure, Cl may be passed over a mixt. of the phosphate with active C. Various details of treatment are described.

Sodium sulfide. I. G. FARBERNIND A-G. Brit. 336,251, June 10, 1929. Anhyd. Na sulfide is made by reducing Na_2SO_4 in the mass with H or other suitable gaseous reducing agent at such a temp. (suitably 500-600°) that the mass is maintained in the solid state throughout the reaction. The reaction may be accelerated by the addn. of a small proportion of Na sulfide or the crude melt obtained in its manuf., NaOH, S, Fe or Ni or mixts. or compds. of these. Cl. C. A. 24, 2253.

Zinc hydrosulfite. HANS PICK. Austrian 120,389, July 15, 1930. In the manuf. of ZnS_2O_4 by treating Zn suspended in LiOH with SO_2 , better yields are obtained by effecting the reaction in the presence of a small quantity of a metal less electropos. than Zn, e. g., Sn or Ag. The addnl. metal may be pptd. in the presence of the Zn, or an inert carrier on which the metal has been pptd. may be used. An example is given.

Zinc oxide. COMPAGNIE FRANÇAISE DE TRANSFORMATION MÉTALLURGIQUE (Repd. H. L. Aubert, inventor). Fr. 693,947, Apr. 15, 1930. In the manuf. of ZnO the greater part of the heat necessary for the distn. of the metal is obtained from the radiant heat of the combustion flame of the Zn vapors. A suitable app. is described.

Enriching natural phosphates. AKTIEBOLAGET KEMISKA PATENTER. Fr. 694,946, May 3, 1930. Natural phosphates are enriched by heating them with H_3PO_4 , so as to form tribasic phosphates of alk. earth metals.

Reactivating carbon, silica gel, etc. JOSEF MICHKA. Austrian 120,098 and 120,101, June 15, 1930. The materials are entrained in a current of a reactivating liquid or gas, in a manner analogous to that employed in cleaning filter sand. App. is shown (120,098). Austrian 120,101 describes modified app. for the same purpose.

Hydrogen and carbon monoxide from methane, etc. I. G. FARBERNIND A-G. Brit. 336,631, April 8, 1929. Gaseous hydrocarbons such as CH_4 or gas mixts. contg. it (other than coke-oven gases or other gases contg. liquid hydrocarbons) are passed, with water vapor (and if desired with further addn. of air or O_2), through a water gas producer in which coke or other carbonaceous fuel is undergoing gasification which is carried out without addn. of CO_2 (the coke having a temp. of at least 800-1000°). As starting material there may be used a natural gas or waste gases from NH_3 synthesis, hydrogenation of coal, cracking, etc. Various details and modifications of procedure are described. Cl. C. A. 25, 783.

Hydrogen and carbon monoxide from methane, etc. I. G. FARBERNIND A-G. Brit. 336,944, April 8, 1929. Various details are described of a process generally similar to that of Brit. 336,635 (preceding abstract).

Hydrogen and oxyacids of phosphorus. I. G. FARBERNIND A-G. Brit. 337,109, Aug. 24, 1929. P or P acids of lower stage of oxidation than H_3PO_4 are heated under pressure with anhyd. H_3PO_4 , e. g., yellow P 10 parts is heated for 3 hrs. with H_3PO_4 85 parts at 300° under 83 atm. pressure, by which an acid contg. 80% P_2O_5 is obtained.

Metallic sodium and potassium. JOSÉ YLLA-CONTE. Fr. 694,587, April 26, 1930. An app. is described for the production of K and Na by reduction of the carbonate or hydroxide with C under reduced pressure produced by pumps drawing off the gases formed. Cooling means is placed between the reaction chamber and the pumps.

Catalysts. I. G. FARBERNIND A-G. Fr. 694,076, April 17, 1930. Very active catalysts are prepd. by disseminating catalytic substances in an extremely fine state of division, preferably in the atomic state, in particularly refractive stabilizers not modified under the conditions of the process and bringing the substance into the state of catalytic activity without using high temps. If oxides or salts are used, they are reduced with active H in the stabilizer.

Polymerization products. I. G. FARBERNIND A-G. Fr. 693,920, April 15, 1930. Polymerization products of good mech. properties are made by heating diolefins, such

as butadiene, while adding alkali metals or their alloys or org compds of alkali metals in portions during the polymerization.

Urea-formaldehyde condensation products. I G FARBENIND A-G (Johannes Kuchenbuch and Karl Eisenmann, inventors) Ger. 516,995, Apr 13, 1929. Added to 490,012 (C A 24, 2256) and 495,790 (C A 24, 3615). Products of reduced inflammability are obtained by effecting the processes of the prior patents in the presence of halogenated aromatic alcs or of di- or polyhydric alcs partly etherified or esterified with one or more halogenated aromatic radicals. Suitable condensation media are *o*-chlorobenzyl alc., glycol monobromophenyl ether and glycol dibromosuccinate. Examples are given.

Plastic masses. DEUTSCHE HYDRIERWERKE A-G Ger. 513,771, Jan 24, 1926. A softening and gelatinizing agent for such plastic masses as cellulose esters, lac, etc., consists of esters of hydroaromatic carboxylic acids which have a wholly or partly hydrated nucleus. These may be combined with other softening agents. Preferably, the agent is an ester of dicarboxylic acids obtained by the oxidation of partly hydrated naphthalene with a wholly hydrated ring, or of hydronaphthols.

Artificial materials. I G FARBENIND A-G Fr. 694,944, May 3, 1930. Artificial materials are made of fibrous materials agglutinated with condensation products of dicarboxylic acids of ethers or thioethers and polyhydric alcs or their solns. The product is submitted to a hardening under pressure. Examples are given of the use of the condensation products of digluolic acid with glycerol in acetone, thiodihydracrylic acid with glycerol in a mixt. of dioxane and acetone, etc.

Composite articles of materials such as rubber and bakelite. B O WARWICK and M WAARWICK (trading as Warwick's Time Stamp Co.) Brit. 336,292, July 12, 1929. Materials such as uncured rubber and bakelite powder are united by molding together under heat and pressure to form a hardened mass.

Pectins and mucilages. HENRI OLIVIER. Fr. 694,460, July 27, 1929. Pectins and mucilages contained in plantain seeds are extd with water in a closed vessel at a temp. above 100°.

Casein. CHARLES C T. PORCHER and JEANNE BRIGANDO Fr. 694,032, April 16, 1930. Casein contg little ash is prepd. by pptn. with lactic acid, AcOH, HCOOH, HCl or H₂SO₄ at the isoelectric pt. of a *pH* value of 4.7. The casein is afterward heated in its serum to 60-65°, the serum decanted and the casein washed in acidulated water. Paracasein and Ca paracaseinate are prepd from com lactic casein.

Dialyzing membrane. FILTRES PHILIPPE Fr. 695,007, Aug 2, 1929. A membrane for dialyzing org products in the presence of solvents is made by dipping a support, e. g., of cloth into a bath of rubber and drying. Fr. 695,008 describes the application of the membranes for the purification of mineral oils.

Porous material for heat insulation, filters, etc. I G FARBENIND A-G Brit. 336,318, July 19, 1929. Siliceous material, such as that obtained in the manuf. of alumina by treating burnt clay with HCl, is heated with alkali such as a 20% NaOH soln and then further heated (sustainably to 300-500°) to effect swelling and production of a porous product suitable for heat, sound or elec. insulation, filters, diaphragms, absorbents, fillers of rubber, lacquer, paper, etc. Numerous examples of siliceous materials and alk. reagents which may be used are given. MgO and MgCl₂ may be employed to strengthen the product, and org fillers may be burnt out to increase the porosity of the product if desired.

Bearing material. CHARLES F NOFTZER U. S. 1,791,834, Feb. 10. Bearings are formed of ground sericite schist and an insol condensation product of a pheolic compd and CH₂O.

Use of diethyl phthalate as a cooling fluid in heat-exchange systems such as engine-cooling systems. DAV J KILLEN (to Waldemar C Wehe) U. S. 1,791,818, Feb. 10. Diethyl phthalate is used alone undild.

Apparatus for calcining seaweed. J C MORRISON. Brit. 336,500, Dec 16, 1929. Structural features.

By-product whitening. JOHN W CHURCH and HARVEY G ELLEDGE (to The Pure Calcium Products Co.) Can. 309,082, Mar 3, 1931. Whitening is produced by the reaction of Ca(OH)₂ and Na₂CO₃, sepn of the major portion of the NaOH formed, formation of a slurry of the impure CaCO₃, treatment of the slurry with CO₂ to convert the remaining NaOH into a carbonate, further treatment of the slurry with CaCl₂ to form NaCl and the washing out of the NaCl.

Applying hood caps to bottles. CARLETON ELLIS (to Ellis-Foster Co.) U. S. 1,792,008, Feb. 10. Binder-carrying caps such as those of heavy paper treated with a mixt. of carnauba wax 80, rosin 18 and S 2% are applied to a container such as a

bottle while maintaining the temp sufficiently high to render the binder of the cap plastic, but low enough not to cause undesirable adhesion between the cap and container.

Use of aqueous oil emulsions to take up shocks in shock-absorbing devices. NOEL B. NEWTON. U S 1,791,842, Feb 10. Emulsions may be used such as those prepd. from water 10 gals, fard oil 2¹/₂ gal, Na_2CO_3 7 lb. and soft soap 2 lb. or from olein 3-5, MeOH 1, mineral oil of 0.845 sp. gr. 19, glycerol 1, KOH 1 and water 75 parts.

Treated receptacle for preventing tarnishing of silverware. GEORGE E. HERRING. U S 1,791,594, Feb 10. Tarnishing is inhibited by treating the inner surfaces of a silverware case with a substance such as Pb acetate which arrests the tarnishing action of H_2S and SO_2 .

Wetting, etc., agents. DEUTSCHE HYDRIERWERKE A-G. Fr 694,230, April 22, 1930. Piperidine and its homologs either alone or with org. acids are used as washing, cleaning, wetting or emulsifying agents. Cf C A. 25, 1048.

"Lime-nitrogen." ELEKTROCHEMISCHE GES. and H. GROMANN. Brit. 336,677, July 20, 1929. Ca carbide is heated in N which is supplied through channels uniformly throughout the mass (as through cardboard tubes in the mass). The mass is heated to 800° or higher and it fuses but remains sufficiently viscous that the channels are not closed.

Apparatus for making lime-nitrogen. GUSTAV HILGER. Ger. 516,445, Sept. 20, 1929.

Device for blowing combustion air into shaft kilns for lime, etc. MASCHINENBAU-ANSTALT HUMBOLDT. Ger. 516,087, Feb 22, 1929.

Rotating furnace for dehydrating gypsum. ETABLISSEMENTS POLIET ET CHAUSSON. Fr 695,029, Aug 6, 1929.

Protecting leather, wallpaper, wood and other materials from fungi. J. G. FARBENIND A-G. Brit. 2,6244, May 1, 1929. The materials (including textile materials other than those of wool) are protected against mildew or other fungi by treatment with a hydroxydiarylmethane or hydroxytriarylmethane contg. one or more univalent halogen phenol residues, e. g., oxide is treated with a caustic alkali soln. of the condensation product from formaldehyde and 2 mol. proportions of *p*-chlorophenol (dil. H_2SO_4 being added to neutralize the alkali). Various other examples are given.

Pencil compositions. EAGLE PENCIL CO. Ger. 516,229, July 16, 1929. Addn. to 505,144 (C A 24, 391). The compn. described in Ger. 505,144 is modified by using other sulfonated oils, e. g., sulfonated olive oil, in place of Turkey red oil.

Stencil sheet. SHUNJIKO HOAN. U S 1,792,003, Feb 10. Sheet material such as Japanese yoshino paper is coated with a compn. comprising a mixt. of esters of polysaccharides such as cellulose acetate, cellulose nitrate, starch acetate, starch palmitate or mannann acetate with naphthene acid glycerides and a heavy hydrocarbon oil. Cf C A 25, 1646.

Materials for electrodynamic microphones. CLINTON R. HANNA (to Westinghouse Elec. & Mfg. Co.). U S 1,792,200, Feb 10. A conical diaphragm has a coil support at its smaller end formed of material such as celluloid to which a coil is united by a cementing material such as a cellulose acetate compn. which when applied contains an ingredient having a solvent action on the material of the support.

Photographic sound records. F. K. CROWTHER and BRITISH TALKING PICTURES, LTD. Brit. 336,579, July 4, 1929. Various mech. details of manual are described.

Dental amalgams. DEUTSCHE GOLD- UND SILBER SCHREIDANSTALT VORMALS ROESSLER (Fritz Ungerer, inventor). Ger. 513,917, May 9, 1928. The amalgams are formed by shaking the constituents in a closed vessel in the presence of a small quantity of dil. NaOH.

Apparatus suitable for heating thermoplastic dental materials, etc. DANIEL S. LOCKWOOD. U S 1,792,434, Feb 10. Structural features.

Dental cement. JOACEN E. THOMSEN. U S 1,792,200, Feb 10. A dry powder suitable for use as a dental cement on addn. of water alone consists of approx. 145 parts by wt. of a silicate cement and 40-60 parts by wt. of pure H_3PO_4 or its equiv. in the P_2O_5 radical, the ingredients being mixed together at a temp. sufficiently low (suitably at room temp.) to retain the mixt. in an unstable condition capable of reacting with the added water to form a plastic mass capable of hardening.

"Frost-preventive" or thawing agents for use on railway switches, etc. G. ERDBRUGGER. Brit. 336,739, Sept. 9, 1929. Ground raw carnallite is used in mixture with liquid bitumen or crude mineral oil.

"Artificial ice" skating rinks. H. REESER. Brit. 336,218, June 4, 1929. A suitable foundation is covered with a layer which may be formed from Na hyposulfite 70

and borax 30 parts with a small quantity of alum, and the surface may be covered with a soap and stearin compn

Mixed felt for use as a box toe material. HENRY P. SHORNECK (to Arden Box Toe Co). U. S. 1,791,537-S, Feb 10 A short-fibered material such as wood pulp 60, is used with leather fibers or wool and cotton fibers 30 and asbestos fibers 10%, felted and impregnated with a thermoplastic material. Cf. C. A. 24, 1478.

Filter for use in tobacco pipes. Z. BRALAY. Brit. 337,200, Dec. 5, 1928 Filter wads are formed of material such as cotton or filter paper which has been treated with dil. H_2SO_4 , heated sufficiently to carbonize the fibers on their outer surface, and then impregnated with $FeCl_3$ or other material such as citric, tartaric or tannic acid adapted to fix the nicotine. The H_2SO_4 still present after the carbonization is neutralized with NH_3 .

Apparatus for supplying fire-extinguishing foam to oil tanks, etc. MINIMAX A-G and PYRENE CO., LTD. Brit. 337,221, Nov. 14, 1929 Structural features.

Composition for fireproofing fibrous materials. S. A. I. G. (Soc. ANON. INVENTORI GUADAGNI) Brit. 336,863, Dec. 10, 1929 Articles are immersed or treated under pressure with a soln. formed of NH_4Cl or $(NH_4)_2CO_3$, 12, H_3BO_3 , 8, NH_3 , 8 and water 70 parts.

Fire-extinguishing compositions. MINIMAX A G. Ger. 516,426, July 28, 1927. Comps. comprising CCl_4 and NH_3 (or a compd thereof) are rendered stable to storage by means of a phys. or chem. drying agent, which may be added to the compn. or located in the container therefor. Suitable drying agents are soda-lime, anhyd. Na_2SO_4 , PrO_3 , and silica gel. Cf. C. A. 25, 1348.

19—GLASS, CLAY PRODUCTS, REFRACTORIES AND ENAMELED METALS

G. E. BARTON, C. H. KERR

Quinquennial review of the mineral production of India for the years 1924 to 1928. Glass-making materials. G. V. HOBSON. *Records Geol. Survey India* 64, 392-9 (1930)—Chem. analyses of Indian glass sands are given (SiO_2 , 90.00, 98.95, 99.39, 98.10 and 99.80).

ALDEN H. EMERY

Dependence of the properties of glass on its thermal history. E. BERGER. *J. Soc. Glass Tech.* 14, 280-94 (1930)—The properties of glass depend upon the state of aggregation within the mass. Annealing produces an inner state of equil. and hence a change in double refraction, density, etc.

H. F. KRIEGER

The decomposition of refractory silicates by fused ammonium fluoride and its application to the determination of silica in glass sands. A. C. SNEAD AND G. FREDERICK SMITH. *J. Am. Chem. Soc.* 53, 483-6 (1931)—The use of an NH_4F fusion for decompn. of glass sands is rapid and quite satisfactory for routine analyses.

M. A. EDDY

Manufacture of photographic lenses. ROBERT S. GULFORD. *Glass Ind.* 12, 23-7 (1931)—The methods employed at the Hawkeye Works of the Eastman Kodak Company are described.

H. F. KRIEGER

Destruction of glass by superheated steam. F. H. ZSCHACKE. *Chem.-Ztg.* 55, 41 (1931), cf. C. A. 24, 4130—No glasses are completely immune to the attack of water vapor, but a typical compn. and list of sources are given for glass which is satisfactorily resistant.

T. H. CHILTON

Term "pounds of glass per pound of coal" is misleading. JOHN W. ROMIG. *Glass Ind.* 12, 28-9 (1931)—The factors requiring consideration in a discussion of melting efficiency are furnace construction, regenerative capacity, temp. in the tank, quality of fuel and pull on the tank.

H. F. KRIEGER

A little-known action of animal glue and other adhesives on glass. GUSTAV GÜNTHER. *Chem.-Ztg.* 55, 37-8 (1931)—A sign was affixed to a store window with dextrin and the glass underneath was "ruined," and damages were sought. It was demonstrated that any adhesive might, when removed after drying, cause an irregular frosting effect, if the glass was old and had been frequently polished. A flexible paste is recommended for such cases.

T. H. CHILTON

Notes on clay and burned clay products. G. S. STAIRS. *Eng. J.* 13, 688-93 (1930)—Some of the properties of clay as a material and the methods of manuf. of such clays which are of importance to structural engineers are indicated. The manuf. of brick and the manuf. of structural tile are outlined. Advantages and disadvantage.

of the car tunnel and the down-draft kilns are listed. A flowsheet of mfg. processes in the clay product industry and a brief discussion are included. W. H. BOYNTON.

Behavior of soluble salts during the burning of clays—the problem of efflorescence. IV. Conditions for the formation of sulfate efflorescences in ceramic materials. A. SIMON AND K. SCHMIDT. *Z. angew. Chem.* 43, 380-2 (1930), cf. *C. A.* 24, 1949. Clay material contg. quartz, or quartz material was admixed with varying quantities of lime (contg. MgO) and heated at 1000° for 3 hrs. and then maintained at this temp. for 5 hrs. in air contg. 0.5 or 2% SO₂. For the same lime content the amt. of sulfate formed rises with increasing SO₂ concn., the amt. of SO₂ absorbed being directly proportional to its concn. Similarly, for the same SO₂ concn. the quantity of this gas absorbed and of sulfate formed increases with the rise in lime content, but at a much lower rate than the latter. The temp. interval 300-600° is most favorable for the absorption of SO₂, max. absorption probably occurring between 400° and 500°. In the firing of ceramic materials, therefore, the interval 300-600° should be passed as quickly as possible in order to avoid absorption of the SO₂ produced by the combustion of the S of the coal. No absorption occurs at 750° and above, the sulfate detected in the expts. carried out at 1000° is formed during the heating up to this temp., and although the dissocn. pressure of CaSO₄ in the presence of SiO₂ is already exceeded at 750°, greater energies than are present are necessary before decompn. occurs. The drying and preheating of ceramic material, therefore, should be completed as far as possible in the absence of SO₂, while for sharp firing, coal contg. S may be used without adverse effect. B. C. A.

Technical analysis of clays with the aid of the microscope. II. Möhn. *Sprechsaal* 62, 731-4 (1929).—The methods employed are described, together with the microscopic characteristics of various clays. B. C. A.

Manufacture of light-weight units of clay and combustible admixtures. A. G. E. BUCH. *Ziegelzeit.* 62, 31-2 (1931).—A fat clay is essential as binder while a variety of admixtures such as sawdust may be used. H. F. K.

Effect of temperature on the cracking of plastic clay; with some notes on Temperature gradients inside the clay during steady heating in a saturated atmosphere and on drying at 100°. S. R. LIND AND P. WHEELER. *Trans. Ceram. Soc. (Eng.)* 29, 314-25 (1930). cf. *C. A.* 24, 5935.—It was shown that clay cracks vigorously when heated in a satd. atm. Hence the failure during drying cannot be due to unequal contractions of different parts of the ware nor to the bursting force developed by entrapped water vapor. A definite relationship exists between the crit. temp. and the water content of the clay in a satd. atm. The actual range over which cracking occurs is a property of the clay. With 4" cubes of clay heated at a regular rate the thermal diffusivity increases as the water decreases from 20.7 to 17.4%, after which it decreases. Clays must be air-dried to a certain water content beyond which the drying process may take place at higher temps. without serious results. H. F. K.

Influence of atmosphere on the load-bearing capacities of fire bricks. II. T. S. SWALLOW. *Trans. Ceram. Soc. (Eng.)* 29, 239-53 (1930).—By use of the Mellor and Moore load test furnace as modified by Dale, several types of bricks were tested in oxidizing, reducing and steam atms. It is suggested that the influence of a reducing atm. at high temps. is apparent more in an increase in the amt. of contraction than in a reduction in refractoriness or load-bearing capacity. At the temps. of these tests (1200-1280°) the action of steam appeared to be negligible as a reducing atm. The Fe content of the brick must be considered when tests are made under oxidizing conditions. It is suggested that the term "oxidizing atmosphere" be defined as that in which FeO cannot exist in stable equl. but is raised to at least Fe₂O₃. H. F. K.

Effect of repeated burning on the structure and properties of lime-bonded silica bricks. II. Determination of the proportions of quartz, cristobalite and tridymite. W. HUGGILL AND W. J. REES. *Trans. Ceram. Soc. (Eng.)* 29, 381-4 (1930); cf. *C. A.* 23, 5289.—Using the optical method of reading the linear intercepts for each constituent in 2 directions at right angles, H. and R. detd. the amts. of the 3 forms of SiO₂ and of glass present after repeated burnings. The amt. of unchanged quartz decreases rapidly in the first 2 burnings and thereafter much more slowly. By extrapolation, it is estd. that at least 20 burnings under these conditions would be necessary to invert the whole of the quartz. Tridymite increases and cristobalite decreases more in the first 6 burnings than later. The quantity of glass remains const. III. Determination of the reversible thermal expansion. *Ibid.* 384-7.—Test pieces were burned up to 12 times in a cone kiln at cone 17, the temp. of the specimens being raised 2.5° per min. The linear expansions for the intervals 15-240° and 15-1000° were, resp. 1 burn 0.73%, 1.152%, 3 burns 0.74%, 1.130%, 6 burns 0.75%, 1.104%; 12 burns 0.49%, 840%. H. F. K.

Some effects of firing temperature, kind of grog and grading on the properties of fire-brick material. A. J. DALN. *Trans. Ceram. Soc.* (Eng.) 29, 320-38 (1930).—A clay giving a refractory test result of cone 28 was used as a bond clay with several highly refractory materials. Ashes of over 60% of fused Al_2O_3 or precalcined diaspore gave good results. Greater rigidity of such materials below 1500° can be obtained by increasing the severity of the kiln treatment and decreasing the grain size of the grog. Crude zirconia gave a very dense vitrified product which, however, squatted under load at relatively low temps. Silica glass and quartzitic grog were about equal, the greater the grain size the greater the after expansion. Vitrified grog should be of finer gradation than the porous grog. When the amt. of clay is not excessive, the value of the clay bonded refractory is governed by the nature of the grog more than by the clay. H. F. K.

Thermal expansion of refractory materials. R. V. WIDMANN, A. JOURDAIN AND H. CASSAN. *Ceramique* 33, 239-56 (1930).—Various types of app are described. Thermal expansion curves are given for silico-aluminous refractories, French, English and German silica refractories, huxite and magnesite. A. J. MOYACK.

The refractory lining of cupolas. JOSEPH RONITSCHKE. *Feuerfest* 6, 113-8 (1930).—Lab tests of refractory brick and examn under actual working conditions indicate that the major factors in the destruction of a cupola lining are chert and much action of the slag. In some cases the life of a lining depends upon resistance to temp. changes. This factor becomes of greater importance the larger the brick. H. A. SOULE.

Electric-furnace production of high-heat-duty refractories. FRED W. SCHROEDER, *Ind. Eng. Chem.* 23, 124 0 (1931).—The operation of the Corhart Refractories Co. is described. Diaspores and laohns are carefully blended, melted in an elec. furnace of the dipping electrode resistance type and poured into sand molds, after which the blocks are annealed for 8-10 days. The resulting product is vitreous, non-porous and refractory up to cone 37, has about 0.5 the linear coeff. of expansion of first class fire brick and is very resistant to the corrosion of glass. H. F. K.

Gas permeability of refractory material, especially at high temperatures. H. IMMER AND W. MUELLER. *Sprechsaal* 64, 85-7, 107-9 (1931).—Two methods for slag gas permeability of refractories at high temps are described. The permeability of 11 pyrometer tubes toward 12 gases detd. at high temps shows that the gas permeability is directly proportional to the pressure and depends on the temp. At very high temps the gas permeability decreases with increased temp. It is also a function of the viscosity of the gas. C. H. LORTON.

The acidification of refractories. V. The relation of flux content and porosity to the acidification of fire brick. HERMANN SALMANG AND OTTO HENNINGER. *Feuerfest* 7, 1-8 (1931).—A mixt. of very pure Al_2O_3 and SiO_2 (1:2) was prepd. Portions were treated with a flux (2% of Fe_2O_3 , CaO , TiO_2 or Na_2O), molded into cupels and fired at 1410°. Six different synthetic slags were made from mixts. of high-grade FeO , Fe_2O_3 , SiO_2 , Al_2O_3 , CaO and Na_2O . The action of each slag on each sort of cupel for 1 hr at 1410° showed that, regardless of the flux, the depth of slag penetration was about the same. The max. was 13-21 mm. for the slag 3.46 CaO 1.46 Al_2O_3 3.63 SiO_2 . The influence of the flux was nullified by the high porosity (53-51 vol %). Crucibles made from a 70:30 mixt. of burnt and unburnt laohn contg. either 0, 2 or 4% flux had a porosity of ≈ 21 vol %. FeO , as a slag, strongly attacked each type of crucible at 1500°. The action was greatest in the presence of 4% Na_2O or TiO_2 . The slag FeO SiO_2 perforated many crucibles. Refractories with 2 or 4% flux were attacked by the various slags only slightly more than those contg. none. Tests on com. brick showed that the action of blast-furnace or open hearth slag is not directly dependent on porosity. Brick of the same chem. compn. and porosity is attacked to the same extent only when the microstructure is the same. H. A. SOULE.

Graphical calculation of the composition of ceramic varnishes (CASTIGLIONI) 26. Firing with powdered coal in German tempering foundries (STOTZ) 9. Structure of silicates (BRAGO) 6. Critical observations on the directions for glass testing of the German Pharmacopoeia, Sixth Edition (KROEMER) 17. Transmission changes in ultra-violet glasses during high-temperature exposure to light (NITSCHE, SCHMUTZ) 3. Heat-treating furnace for enameled articles (U. S. pat. 1,702,281) 9. Alloys for sealing to glass (Brit. pat. 337,083) 9. The ternary system: K_2O - CaO - SiO_2 (MORAY, et al.) 2.

Tonindustrie-Kalender, 1931, 2 vols. Berhn: Tonind.-Ztg. M. 4. Reviewed in *Ceram. Abstracts* 10, 236 (1931).

WIDEMANN, R. V., JOURDAIN, A., AND CASSAN, H. - L'étude de la dilatation des matériaux réfractaires. Paris La Ceramique 20 pp. Reviewed in *Ceram Abstracts* 10, 236(1931)

Paddle mechanism for circulating molten glass. ALEXANDER F. McNISH (one half to John Moncrieff, Ltd.) U S 1,792,286, Feb 10. Structural features.

Apparatus for supplying molten glass to a suction-gathering machine. J. MONCRIEFF, LTD., and A. F. McNISH. Brit. 335,883, Aug 9, 1929. Structural features.

Device for gathering molten glass and similar materials. EGYENCLT IZZOLAMPA ES VILLAMOSSAGI RESZEVNYTARSASAG and J. LEVAI. Brit. 335,931, July 2, 1929. Mech. features.

Apparatus and method for circulating and feeding mold charges of molten glass. KARL E. PERLER (to Hartford Empire Co.) U S 1,791,634-5-6, Feb 10. Mech. features.

Forming tubes or rods of glass. P. A. FAIR. Brit. 336,460, Nov. 14, 1928. App. is described.

Apparatus for glass bottle manufacture. CHARLES BADGER (to Lynch Glass Machine Co.) U S 1,792,267, Feb 10. Structural features.

Coloring glass. DEUTSCHE GASGLÜHLICHT ALER GES. M. B. H. Fr. 694,828, May 1, 1930. Glasses are shaded in red with oxides of rare earth metals while preserving the dichroism by adding small quantities of colloidal colors known to color glass red, Au, Cu O or Se. Cf. C. A. 24, 5471.

Tempering glass. SOC. ANON. DES MANUFACT. DES GLACES ET PRODUITS CHIM. DE SAINT-GOBAIN, CHAUNY ET CIREY. Fr. 695,035, Aug 7, 1929. An arrangement is described for blowing air on to the heated glass. Fr. 695,040, Aug 8, 1929. An app. is described for tempering glass. It consists of 2 hollow plates through which cooling liquid flows and between which the sheets of glass are mounted. Cf. C. A. 25, 571-2, 790.

Glass sheet. Y. BRANCART. Brit. 336,921, Feb 14, 1929. Mech. features relating to the formation of the sheets (or their immediate treatment after formation) with a rotating rough roller run at such a speed that the glass is driven backward slightly toward the beginning of the plate. The impressions formed disappear at once. Brit. 336,924 relates to a generally similar process including rolling with smooth rollers. Cf. C. A. 25, 571.

Table for receiving or casting glass plates or sheets. C. LOCREILLE. Brit. 336,550, April 9, 1929. Casting tables are made of a sandless concrete composed of fragments of porphyry and of a cement formed of a ground mixt. of burned lime, silica and alumina (the porphyry fragments being distributed in several layers through the thickness of the table in graded sizes with pulverized porphyry in the top layer).

Apparatus for making sheets of glass. MISSISSIPPI GLASS CO. Fr. 693,940, April 15, 1930.

Apparatus for rolling sheets of glass. N. V. MAATSCHAPPIJ TOT BEHEER EN EXPLOITATIE VAN OCTROOIEN. Fr. 694,973, Apr 4, 1929.

Rolling and fire-polishing glass sheets. A. BRANCART. Brit. 336,533, Feb 23, 1929. App. and various details of procedure are described.

Plate glass rolling machine. N. V. MAATSCHAPPIJ TOT BEHEER EN EXPLOITATIE VAN OCTROOIEN. Ger. 513,879, Nov 29, 1929.

Sole plate for glass-annealing furnaces, etc. YVON BRANCART. U S. 1,791,656, Feb 10.

Furnace for annealing glassware. W. A. MORTON. Brit. 336,596, April 15, 1929.

Apparatus for making glass reinforced with wire. MISSISSIPPI GLASS CO. Fr. 694,215, April 19, 1930.

Safety glass. JOHN NEWTON. Fr. 693,919, April 15, 1930. Safety glass is made by incorporating 2 or more sheets of celluloid or cellulose acetate, etc., between 2 sheets of glass.

"Unsplinterable glass" sheets. T. W. HOLT and J. P. W. STUART. Brit. 336,037, Aug 1, 1929. Various details of app. and procedure are described for uniting the component sheets with use of fluid pressure in an autoclave.

Uniting glass sheets with non-brittle material. LEBEY-OWENS GLASS CO. Brit. 336,732, Aug 30, 1929. In uniting glass sheets with cellulose derivs. or synthetic resins or with a combination of the two, the latter is sprayed with a bonding medium such as a mixt. of di Bu phthalate and a polyglycol deriv. such as the monoethyl ether of diethylene glycol, and the sheets are united by heat and pressure. The glass may be sprayed with a cellulose deriv. compn. or synthetic resin compn. in a low b. p. solvent.

mixed with higher b. p. solvent or plasticizing substances. Various details of procedure and compns are given.

Uniting glass sheets with sheets of cellulose derivatives. S. BARRATT. Brit. 335,890, June 29, 1929. For removal of knife marks and the like, a cellulose deriv. sheet is softened and pressed with a hard surface (which may be a gelatin-coated glass sheet with which it is to be united). Various details of heating and pressing the sheets together are described.

Joining glass sheets with an intervening sheet of celluloid or similar materials. T. W. HOLT and J. F. W. STUART. Brit. 337,158, Oct. 2, 1929. Various mech. details for uniting the sheets are described.

Composite sheets of glass and cellulose derivatives. BRITISH CELANESE, LTD. Brit. 335,878, March 13, 1929. Two glass sheets are united by an intermediate sheet formed of a cellulose deriv. and an org. tartrate such as di Bu or di Lt tartrate and dyes, synthetic resins and stabilizers such as urea also may be added. The cellulose deriv. sheet may be united to the glass by an adhesive such as gelatin or vinyl acetate resin or a suitable mixt. of a cellulose deriv. and synthetic resin.

Fused silica sheets suitable for window panes. J. H. SULLIVAN (to British Thomson-Houston Co., Ltd.). Brit. 336,923, Feb. 14, 1929. Fragments of cryst. SiO_2 are fused *in vacuo*, the material while plastic is allowed to assume, by gravity, the form of a flat sheet, and is caused to harden under gaseous pressure, as by admitting air or by use of N under pressure (suitably 200 lb per sq in). App. and various details of procedure are described.

Purifying clay. WM. FELDENEHEIMER. U. S. 1,791,950, Feb. 10. A liquid suspension of the material is deflocculated, impurities are sep'd (as by sedimentation) and the slip is flocculated and thickened with reagents such as lime and HCl in quantities which render the clay non acidic, and the clay is sep'd from the liquid.

Refining kaolin and clay. ERICH MÖLLER. Ger. 516,144, Dec. 30, 1926. Water and water glass in const. proportions are continuously fed to a mixing vessel to which kaolin or clay is also fed at such a rate that a suspension having a sp. gr. between 1.08 and 1.09 is produced. It is stated that a streaky appearance develops on the surface of the suspension when the sp. gr. lies between the limits mentioned. The suspension flowing from the mixing vessel is passed upwardly through a funnel shaped vessel to remove fine impurities, and then neutralized with HCl and worked up in known manner. Cf. C. A. 25, 783, 790.

Porous bricks. T. KOMT and F. GORLICH. Brit. 337,231, Nov. 20, 1928. An extrusion press is provided with means for injecting compressed air into the plastic material for forming the bricks before the material is fed to the extrusion nozzle.

Heat-insulating bricks, etc. STERCHMOLWERKE G. M. B. H. Ger. 516,377, June 8, 1926. Porous bricks, etc., are evacuated and then coated with a material impermeable to gas, e. g., rubber. The coating may be partly applied before evacuation.

Rotary ceramic furnace. GEORG VON HERTZKA. Ger. 513,583, May 12, 1929. The bearing stones are partly of high Al_2O_3 content, and partly of chamotte stone.

Ceramic kiln. JEGOR I. BROWN. Ger. 513,581, July 28, 1926. An electrically heated tunnel kiln for firing ceramic ware at high temps. and with pre heating, firing and cooling zones is described.

Burning ceramic sludge. LEOPOL, INTERNATIONALE PATENTVERWERTUNGS G. M. B. H. and N. V. "SOLOPOL" INGENIEURSBUREAU TOT EXPLOITATIE VAN HET, SYSTEM POLYVUS. Ger. 513,584, Sept. 7, 1927. Ceramic sludge, especially crude cement sludge, is wholly or partly dried, heated in a rotary furnace until granulated, calcined and finally sintered.

Molding fused quartz. DEUTSCH-ENGLISCHE QUARZSCHMELZE GES. Brit. 336,403, Dec. 12, 1928. In forming articles such as insulators or tubes with ribs or flanges, a hollow blank of the material is expanded in a mold by compressed air and then further shaped by application of pressure mechanically.

Apparatus and hydraulic classification system for grading abrasives such as those used in continuous glass-grinding apparatus. PILKINGTON BROS., LTD., F. B. WALDRON and J. H. GRIFFIN. Brit. 336,587-8, July 12, 1929.

Abrasive wheel. JOHN R. GAMMETER (to George W. Perks Co.). U. S. 1,792,083, Feb. 10. An abrasive wheel body is formed with spiral grooves in its opposite faces, and wire convolutions are wound while hot in the grooves and permitted to shrink in them, and over these a bonding layer of material such as a rubber or synthetic resin compn. is cured under pressure to secure the wire convolutions to the face of the wheel.

Compositions for grinding-wheels. CARBORUNDUM CO. Brit. 336,631, July 17,

1929 A cement formed of S 95 and ground coke 5% is used in building up wheels of segmental abrasive blocks (various structural details of which are given).

Refractory earth. FEHLNER & ZIEGLER A-G. Fr. 694,818, May 1, 1930 Refractory earth is baked in a rotating tubular furnace at temps. of 1100-1300° and the furnace is preferably enlarged in the 800-1000° zone

"Semi-porcelain" composition. J. E. TAMS. Brit. 336,172, Dec. 27, 1929 Asbestos powder 50, ball clay 40 and either ground stone 20 or feldspar 20 (or of each 10) parts are blunged with water into a slip or plastic mass for treatment in the usual way

Enamels. W. E. DOUGHERTY. Brit. 336,797, Oct. 17, 1929 TiO_2 is used as an opacifier in enamels which do not contain F and CaCO_3 or other suitable salt of Ca , Ba , Sr or Mg is preferably added to the batch to intensify the opacity produced by the Ti Small quantities of Zr or Sn oxides may be added to counteract any tendency to yellowish color Various details are given

Coloring enamel, glass, etc. DEUTSCHE GLASFABRIK-ANTER-GES. M. B. H. Fr. 691,884 May 2, 1930 To obtain red colored enamel, glass or glazes with Se , use is made of enamel, etc., which has a high content of rare earths Cf. C. A. 24:5451.

20—CEMENT AND OTHER BUILDING MATERIALS

J. C. WITT

Quinquennial review of the mineral production of India for the years 1924 to 1928. Cement. A. M. HESON. *Records Geol. Survey India* 64, 359-70(1930) —The Indian cement industry has increased materially each yr. Several companies use a limestone which naturally contains the correct proportions of clay. Building materials. E. H. PASCOE. *Ibid* 347-59 Marble. *Ibid* 406-9 —Marbles of all colors are plentiful in India Slate. E. L. G. CLEGG. *Ibid* 429-31 ALDEN H. EMERY

Electro-melted cement as a new building material in the chemical industry. CONSTANTIN REDZICI. *Apparatsbau* 42, 277-8(1930) —The manuf. of a cement from bauxite and burned lime in an elec. furnace is described The product has the following approx. compn: Al_2O_3 35-50, CaO 35-45, Fe_2O_3 5-15, SiO_2 5-10% It is compared to portland cement and may supplant it M. C. ROGERS

Blast-furnace-slag portland cement. GIUSEPPE TOMARELLO. *Atti III congresso naz. chim. pura applicata* 1930, 557-60 —Compn. and properties of various slag cements are reviewed E. M. SYMMES

Potash salt recovery in cement burning, with special reference to Estonian potash ores. JAAK KUVSK. *Acta Commentationes Univ. Tartuensis (Dorpatensis)* AXA, No. 4, 1-36(1930)(in German) —Fathonia contains relatively large amts. of K ores capable of being burned to cement. In burning, K compds. are volatilized in part and can be pptd. electrostatically Since the K compds. in different cement plants volatilize quite differently, and since the factors affecting the volatility are not known accurately, volatilization of K compds. was investigated in the lab. In a mixt. of K silicate with CaCO_3 , in the ratio $2\text{CaO}:\text{SiO}_2$, decompn. of K silicate and volatilization of K began simultaneously at about 1150° With increasing temp. and duration of heating the amt. of K volatilized increased per unit of time. If less CaCO_3 were added to the K silicate than corresponds to Ca metasilicate, volatilization of K stops almost wholly, even at very high temp. Starting from these mixts., the volatility of K increases rapidly with increase in amt. of CaCO_3 added Addn. of Fe_2O_3 affects K volatility very favorably because it increases the velocity of K silicate decompn. at low temp. The Fe_2O_3 added remains almost unattacked Addn. of Al_2O_3 is unfavorable to K volatility because most of the CaO present changes to aluminate Addn. of CaSO_4 is just as effective as addn. of CaCO_3 In K volatilization from synthetic K aluminate the same general results were obtained The amt. of K volatilized at 1440° was low if only sufficient CaCO_3 were added to the aluminate to form only Ca meta aluminate $(\text{Ca}$ or $\text{K})_2\text{SiO}_3$ and $(\text{Ca}$ or $\text{K})_2\text{Al}_2\text{O}_3$ are unstable at the roasting temp. in question, while $(\text{Ca}_2$ or $\text{K}_2)\text{SiO}_4$ and $(\text{Ca}_2$ or $\text{K}_2)\text{Al}_2\text{O}_4$ are unstable and lose K_2O F. M. SYMMES

Hydrated lime as used in engineering. R. J. HOLDEN. *Eng. News-Record* 106, 148-9(1931) —Curves are given showing the total annual production of CaO and Ca(OH)_2 during the period 1906-1928 and the amts. used for building during that period, together with a brief discussion of the properties of CaO and Ca(OH)_2 , with particular reference to deterioration on storage There has been no increase in the production of CaO since 1906, but the production of Ca(OH)_2 has markedly and steadily increased

R. E. THOMPSON

A study of methods of curing concrete pavements. Report of field investigation in Tennessee. F. H. JACKSON AND E. W. BAUMAN. *Public Roads* 11, 209-35 (1931).—The method of cure involving the application of wet burlap for periods varying from 24 to 96 hrs without further curing compares favorably with the standard method. The application of Sinalkraft paper for periods of 24 to 48 hrs without further curing is a reasonably satisfactory method of cure as compared with the standard method. In the poor earth cure it is not necessary to keep the earth saturated with water for the entire 10-day curing period. The sodium silicate method of cure gives results somewhat less satisfactory than those given by the standard method. Surface application of CaCl_2 appears to give results nearly as satisfactory as the standard method. Curing with bituminous materials, as used in these tests, was unsatisfactory. There was marked evidence of increased transverse cracking. Nothing is gained by the use of tar paper on the subgrade.

I. B. MILLER

Revision of "Directions for testing (D I N 1995) and delivery (D I N 1996) of road and construction materials containing asphalt and tar." ANON. *Asphalt u. Teer* 30, 1253-62 (1930).

K. H. ENGEL

Testing limestone for road building. J. JACHTEL. *Asphalt u. Teer* 30, 1011-2 (1930).—Sols in CO_2 -contg. water (indicative of grain size) and mech. hardness should be used. The presence of $\text{Ca}(\text{HCO}_3)_2$ is a proof of harmful weathering.

K. H. I.

American road tars. HUGO LABER. *Brennstoff Chem* 12, 80-90 (1931).—Comparison of grades, consumptions and prices with corresponding German products.

F. W. JUNG

German patent development in bitumen emulsions for road construction. A. VON SKOPNIK. *Teer u. Bitumen* 28, 581-5 (1930).

K. H. ENGEL

Road asphalt and mechanical methods for testing it. W. BIRKHÄUSER. *Asphalt u. Teer* 30, 911-46 (1930).

K. H. ENGEL

Preparation of a new road binder. C. P. BROADHEAD. *Inst. Gas Eng., Communication* No. 3, 23 pp (1930).—See C. A. 24, 4372.

G. G.

Determination of wool content in roofing felts by chemical methods. R. KORN AND B. SCHULZ. *Teer u. Bitumen* 29, 31-6 (1931).—Admixed wood flour is insufficiently dissolved during treatment of the felt with 60% H_2SO_4 . Shaking during the digestion reduces the residue to 6%. A felt sample equiv. to 10 g. dry, ash free material is boiled in H_2O and macerated. The fibers, isolated and washed with alc., are shaken with 300 cc. H_2SO_4 (80%) for 3 hrs. The ash is removed with H_2O and dil. NH_3 , and the residue dried at 100-5°. Ash anal. If warranted, 6% of the wood flour content are deducted from the wool residue. The wood flour of the felt should be extd. microscopically in mnges of 10%.

K. H. ENGEL

New methods of wood utilization and wood research. HANS WINTERGUS. *Z. Ver. deut. Ing.* 74, 1169-78 (1930).—A short outline of wood chemistry is followed by details of kiln drying, impregnation, wood decomposition processes, surface clarring to enhance the grain effect, liquid infusion of living trees, turpentine recovery, wood cellulose, saccharification, artificial silk and metallized wood.

H. M. SYMES

Protecting wood from fungi (Brit. pat. 336,214) 18.

Cement. F. M. L. REUBKE (to I. G. Farben und A. G.). U. S. 1,791,681, Feb. 10. See I. R. 662,145 (C. A. 24, 482).

Cement. TREVOR M. CAVEN (to Frederick B. Hazard). U. S. 1,791,691, Feb. 10. A cement in dry powd. form suitable for monolithic linings or walls comprises about 5% silica in colloidal form, about 10% silica in diatomaceous form and about 85% silica in relatively coarse form, all these forms of silica being of approx. the same content.

Dolomite cement. GUSTAV MATSCHAK. Ger. 600,138, Mar. 25, 1928. MgSO_4 is added to the product in order to react with any CaO formed during the firing and at the same time to raise the Mg content.

Portland cement. JAMES H. COLTON (to Pacific Portland Cement Co.). U. S. 1,791,800, Feb. 10. A portland cement mix is burned to form a clinker, the clinker is ground, the ground clinker is added to a raw cement mix to form a granular mix and this mix is burned to form a clinker.

Waterproof cement. S. A. I. G. (Soc. ANON. INVENZIONI GUADAIGNIN). Brit. 336,860, Dec. 9, 1929. See I. R. 680,193 (C. A. 25, 560).

Composition for waterproofing cement. SAM C. BENNETT. U. S. 1,791,704, Feb. 10. A mix formed from Na_2CO_3 , 9, $\text{Al}_2(\text{SO}_4)_3$, 1, KMnO_4 , 0.03 and water about 20 parts is used together with pure cement in sufficient quantity to leave the compn. fluid so that

it may be applied as a liquid coating to waterproof and "rejuvenate" worn and discolored cement.

Rotary kiln for the preparation of fusible cement. ARNO ANDREAS Ger 513,582, Feb 7, 1929

Apparatus for granulating crude material and burning to produce cement. OTTO LELLER Ger 515,172, Feb 10, 1929 Addn to 466,298 (C. A. 23, 497).

Hydraulic product. KARL BIRHL Fr 695,070, May 5, 1930 A hydraulic product for addn to lime, cement, etc., is made by heating all sorts of clay, marl products of volcanic eruption, etc., and fine grinding

Concrete. LUDWIG HEYNEMANN Fr 695,099, May 5, 1930 A concrete which is resistant to sea water is made by adding to ordinary or superior cements or mixts of these with sand, etc., finely divided cinders from the combustion of powd coal.

Concrete composition. CARLTON E MILLER U S 1,701,630, Feb 10 A concrete which sets quickly to a high strength is formed from an admixt. of portland cement with a preformed mixt. comprising CaCl_2 2 and diatomaceous earth 3 parts (suitably 5% as much of this mixt. as of portland cement)

Cellular concrete. TORRENT-GRS M. B H Fr 694,287, April 22, 1930. A liquid absorbing substance such as kieselguhr is added to cellular concrete so that the mass is plastic without losing its swelling power

Cellular or porous concrete. EMMIG I LINDMAN (to Aerocrete Corp. of America), U S 1,791,820, Feb 10 Cement and water are mixed with a metal powder which is adapted to cause generation of gas and with slag so divided as to present fine hydrating particles and larger particles of a size and in sufficient quantity to prevent formation of cracks during setting (the slag not exceeding 80% of the wt. of the total dry ingredients of the concrete, and at least 10% having a grain size less than three-eighths of an in.).

Casting concrete pipes. ABRAHAM RUTENBERG Ger 516,437, May 17, 1929 See Brit. 318,285 (C. A. 24, 2266)

Dust-lysing agent for roads, etc. FRITZ LYDTH Ger. 513,846, May 15, 1925 Low temp tar (e.g., generator tar, crude mineral tar, etc.) from brown or stieic coal is used, alone or mixed with other residues or distn products

Asphalt or tar compositions. J G FARBENVIND A-G Fr. 694,789, May 1, 1930 Asphalt, tar or bitumen compos for roads, paint, etc., contain appropriate polymerization products of diolefins.

Colored mineral granules suitable for decorating roofing shingles. RALPH L. ATKINSON (to Arthur D Little, Inc) U S 1,702,048, Feb 10 Material such as compacted raw clay or shale is reduced to granules of the desired size, these are send from material of finer size or dust, treated with water or a coloring agent such as Fe, Cu, Mn, Co or Cr compds if required, and heated to a hardening temp without fusion.

Artificial stone, etc. OSTERREICHISCH-AMERIKANISCHE MAGNESIT A-G and KONRAD LEDMANN Austrian 120,418, July 15, 1930 Addn to 103,151 In the manuf. of shaped articles from fibrous materials and a Sorel cement, the compn is bound and hardened in the mold by means of a gas heated to a temp between 220° and 450°. Open, closed or perforated molds may be used, and the gaseous atm. may contain CO_2 . Cf C. A. 24, 4913

Mold for casting plates from a mixture of magnesite, magnesium chloride solution and powdered stone or quartz. MICHAEL GUTTENBERGER. Ger 516,066, Aug 23, 1927

Molded material resembling wood, etc. A. HAWESLANDER (to Halizite Corp.) Brit 336,754, March 13, 1929 Fibrous material such as sawdust is first treated with a small quantity of a volatile liquid such as "petrol," is then incorporated with sufficient phenol CH_3O resin condensation product to cover its surface and is molded under heat and pressure so as to permit venting of the volatile liquid before the binding agent hardens. The material may be molded in contact with surface veneers of wood. Various details of procedure are described

Preservative for wood. HOLZINDUSTRIE-WERKE JOSEF BENZ A-G Ger 516,407, Dec. 14, 1929 Asphalt rubber melts are dissolved in a mixt. of acetone oil with homologs of benzene and (or) with hydronaphthalenes

Preserving wood. LEO PATRICK CURTIV. Ger 518,075, Mar 12, 1927. See Brit. 291,857 (C. A. 23, 1241).

Impregnating wood. IMPREVA HOLZIMPRÄGNIERUNG UND HOLZVERWERTUNG A-G Ger 513,848, Jan 16, 1925 The bottoms of masts are treated with a soln. of bitumen, pitch, etc., in CCl_4 .

Apparatus for impregnating wood with creosote and for compressed air or steam treatment. C GOODAY Brit 227,133, Sept. 18, 1929 Structural features.

Apparatus for the partial impregnation of wooden masts or poles. ERNST REICH and LUDWIG TRAMER. *Austrian* 120,412, July 15, 1930.

21—FUELS, GAS, TAR AND COKE

A C FIELDNER AND ALDEN H EMERY

Quinquennial review of the mineral production of India for the years 1924 to 1928. Coal. CYRIL S FOX. *Records Geol. Survey India* 64, 31-78 (1930).—Complete statistical data for production and handling of Indian coal from 1919 to 1928 are given. Typical coal analyses are given. Coal reserves are estimated. ALDEN H EMERY

The evolution of coal. HENRY BRIGGS. *Chemistry and Industry* 50, 127-33 (1931).—Coal of all varieties has undergone and is still undergoing a species of slow spontaneous change involving the discharge of CH_4 and H_2O . It calls this auto metamorphism. It may proceed entirely independently of outside influences, although heat (mild) and pressure accelerate, and certain influences, such as a large quantity of mineral matter, retard the action. The continued evolution of CH_4 from coal is confirmatory proof of this reaction. Analyses of coal of all ranks, ash, S- and N free, are plotted on a C-H-O triangular diagram. The gaseous evolution necessary to move from one rank of coal to another is called. ALDEN H EMERY

Bibliography of United States Bureau of Mines investigations on coal and its products, 1910-1930. A C FIELDNER and M W VON BERNEWITZ. *Bur. Mines, Tech. Paper* 493, 56 pp (1931).—Over 750 references to articles on coal and coal products by members of the staff of the U S Bureau of Mines are classified and cross indexed. The subjects covered include origin, constitution, analyses and classification of coal; chemical and physical properties of coal and coke; absorption and evolution of gases in coal; corrosion and mine water, coal dust inflammation, methods of sampling, analyzing and testing, mining, coal preparation, storage and spontaneous heating, processing, including briquetting, high temperature and low-temperature carbonization, utilization, including combustion in power boilers, domestic furnaces, locomotives, marine boilers, kilns and oil stills, smoke abatement, lignite; and economics and statistics. ALDEN H EMERY

The primary oxidation of bituminous coal. W FUCHS and O HORN. *Brennstoff-Chem.* 12, 65-7 (1931), cf *C A* 23, 1403, 24, 4767, 25, 398.—For conversion into humic acids, the humins of bituminous coals have been readily oxidized in a short treatment with HNO_3 (sp. gr. 1.40). The products possess the characteristics of hydroxy carboxylic acids, form permittite like K salts and ether-esters. They are partly sol., especially when damp, in organic solvents, as pyridine, acetone, "methylglycol," but are not always peptizable in alkalies. The peptized acids are identical with the original except for a difference in soly (also produced by a H_2O pyridine treatment) and a 20% greater K content in their K salts. The humin content of striped (bituminous) but not anthracite or boghead coal may be made water-sol. in this way. F W JUNO

The present status of low-temperature carbonization technique. CH. BERTHELOT. *Rev. metal.* 27, 672-9 (1930).—From the standpoint of its present development in England and Central Europe low temperature carbonization is faced with 3 major difficulties: (1) high cost of production, which in some cases is 3 times that of high-temperature coke, (2) little profitable disposal of the primary tars in other than the fuel markets, and (3) the competition following new methods of utilizing high temperature carbonization gas. Hydrogenation and cracking methods are promising for increasing the value of the tars. H L OLIN

Preparation of fiber coal (fusain) through artificial carbonization. TH. LANGE and P. ERASMUS. *Braunkohle* 29, 463-9 (1930).—Samples of silicified lignite with well preserved cell structure (trunk and roots) were carbonized by heating under water at 340° for 2 hrs. SiO_2 was leached out with H_2F_2 . Microscopic slides showed that the organic material had been carbonized to fusain. Isolated cellular aggregates, filled with bituminous material, had changed to vitrain. Conclusion: Cell walls filled with humic acids were gradually resorbed and changed to vitrain. Wood which was filled with gases, SiO_2 , etc., or which had lost its permeability to humic acids through drying, retained its structure and was carbonized to fusain. K H ENGEL

Relation between the contents of coal in volatile matter and in ash (supplementary note). M. LEGRAYE. *Ann. soc. geol. Belg.* 53, B55-7 (1930).—Replying to a criticism of Denoel of a previous paper (*C A* 24, 710), L. gives more curves and discussion. He finally concludes that there is no relation between ash content and content in volatile matter, and that no conclusions can be drawn regarding the comparative origin of bituminous and anthracite coals. R. H. EWELL

The relation between gross and net heating values and volatile matter of coal. WILHELM GUMZ *Feuerungstechnik* 19, 1-3(1931).—From 100 analyses in the literature G derives the formulas $Q = 8150 + 6543V - 17,308V^2$ and $H = 8150 + 4489V - 13,804V^2$ for the gross and net heating values, resp., of the dry ash free coal V being the fraction of this which is volatile.

Diffusion as a factor in burning pulverized coal. DOUGLAS HENDERSON *Power* 73, 238-9(1931).—While the mixing of air with a gaseous fuel is in part by diffusion, macroscopic fuel particles must be mixed mechanically with air. D. B. DILL.

Testing coking coals. WOLFGANG MELZER. *Arch. Eisenhüttenw.* 4, 169-75 (1930).—A no. of tests are outlined for Westphalian coking coals. For works labs., extn methods are too time-consuming, and sufficient information as to the coking qualities can be obtained by detn of the softening point. H. STROBAT.

Asturian coals. III. Rejuvenation of a gas coal. BENITO A. BUYLLA. *Anales soc. españ. fis. quim.* 28, 859-66(1930), cf *C. A.* 24, 2860.—In "berginizing" a coal having volatile matter 32% (yielding highly swelled coke 66.38 and primary tar 5.8%), 22.2% oils and 64% solid residue were obtained. The residue contained volatile matter 28.94% and produced hard, coherent coke 69.54 and primary tar 9.41%. This indicates the possibility of rejuvenating coal by hydrogenation. E. M. SYMES.

Coal as a raw material. E. P. ARMSTRONG *Chem. Markets* 28, 158-60(1931).—A plea is made for the most efficient means of utilizing coal, coke, oil, etc. Hydrogenation after carbonization of the coal under conditions assuring a max. yield of oil and a char which may be used in the manuf. of water gas and H₂, appears economically best.

Humic acids from brown coal. Preliminary communication. JANS STACH. *Z. anorg. Chem.* 44, 118-20(1931).—Not only humic acid but also brown coal from which the bitumen has been extd. can be methylated. The ability to react with diazomethane proves the acid character of the extd. coal. A distinction is made between ortho- and meta-humic acids the latter not being sol. in cold KOH. E. SCHOTTE.

Recent developments in by-products from bituminous coal. A. C. FIELDNER. *Bur. Mines, Rept. of Investigations* 3079, 13 pp (1931).—There has been steady gain in value and yield of by-products from coal during the last 15 yrs., the design of ovens and by-product recovery equipment has also shown const. improvement. The by-product oven has recently supplied much base load gas for city use. Natural gas and refinery gas are increasingly important competitors. Cheap synthetic NH₃ has made the profitable recovery of by-product NH₃, a problem. The Tern and C. A. S. processes for the conjoint recovery of S, H₂, and cyanogen, European developments and the recovery of by-product S peculiarly adapted to fungicidal and agricultural purposes as practiced in America are described. The solution of the phenol recovery problem and the reduction of light-oil losses by improvements in refining are discussed. ALDEN H. FINEAY.

Concretions in coal. X. STALNIER. *Ann. soc. sci. Bruxelles* 50, Ser. D, 247-55 (1930).—The formations known as "coal apples" found in English pit-coals and in certain anthracite coals of the U. S., nodules found in Australia, ball coal in pit-coal of Hindustan and certain concretions found in Belgium are described. S states that no decisive proof of the origin of these structures is available from these studies. The phenomenon is probably more complex than one thinks and may be due to a combination of causes. Even the term concretions is used for want of a better word. They are not true nodules, produced by concn. at a given point, of substances at first scattered through a rock. Various possible theories are discussed. ALICE W. EPPERSON.

Coal-cleaning problems. L. W. NEEDHAM. *Iron Coal Trades Rev.* 121, 525-7 (1930).—A review. LADSLIA B. BRAGO.

The hydrogenation of solid fuels. CARL ZERRE. *Chem.-Ztg.* 55, 4, 18-9, 38-40, 94-6, 114-7, 136-7, 152-3(1931).—A review of the present status of the process, with particular reference to recent patents, and extensive bibliography. T. H. CHILTON.

Fusion points of the ash constituents of the Sonnenschein seam. H. WINTER AND H. MÖNNIG. *Glückauf* 67, 156-60(1931).—Fusion points of ash from different structural parts of the Sonnenschein seam detd. by the methods of Bunte and Baum (*C. A.* 22, 1457) and of Dolch and Pochmüller were fusain 1000-1150°, vitrain 1125-1200° and durain 1530°. In every case as the compn. of the ash approached that of pure clay, Al₂O₃, 2SiO₂, 2H₂O, the m. p. was raised. H. I. OLIN.

New researches on ash fusion points. C. STAEMMLER. *Chem.-Ztg.* 55, 59-60 (1931).—Various Ca, Fe and Si salts were combined into mixts. and the fusion points detd. by thermocouples after the method of Bunte and Baum (*C. A.* 22, 1457). Sulfates always raised the fusion points of the mixts., probably because of some interaction between S and the oxides. C. R. FELLERS.

Furnace heating by means of pulverized fuel. G. E. K. BLYTHE. *Metal Ind.* (London) 34, 177-9, 231-2 (1930). R. J. C.

At what temperature should purification masses be dried? Dr. VOOER. *Ind. Gas* 50, 501-5 (1930). Expts. on drying of $\text{Fe}(\text{OH})_3$ were conducted to det. the correct drying temp. for purific. masses. Even at 110° not all water is removed after drying for a long time because of the colloidal character of the hydrate and the lack of a definite hydrate. This holds for 1st mass as well as for the weathered limonite which is used as a natural product. Every drying temp. is arbitrary. The method outlined is as follows: Weigh at least 1 kg. of substance in a flat dish and place in a drying furnace at $60-70^\circ$. After 2-3 hrs. take out the dish and expose to air for 15-16 hrs. (overnight). Repeat the cycle of heating, aeration and weighing (usually three) until const. wt. (within 1%) is obtained. The analyses for Fe of the mass should be related to the original wt. and not calcd. on dry wt. This is also true for the absorption capacity detn. for H_2S (generally 74% by wt. of the Fe present, 10% of the mass itself). R. J. C. VAN DER HORVEN.

Alcohol as motor fuel. Report on practical large-scale tests. F. W. FULMANN. *Chem. Ztg.* 54, 518 (1930). A suitable substitute for gasoline used in Brazil in 1916 consisted of a mixt. of alc., EtOH , Et_2O and petroleum in the wt. proportions 73:15:10. R. L. SCHAFER.

Alcohol production for motors in Brazil. FREDERICO W. FREIRE. *Ind. Supr. J.* 32, 404 (1930). In recent years much has been done on the utilization of molasses for the manuf. of motor spirits. One plant with a capacity of 100 tons of motor alc. is operating at its full capacity. In a test of "Asulina," a national motor spirit containing 57% ether, on a motor car fitted with 2 T. A. G. motors of 100 h. p. each, av. fuel consumption was 1.4 l. "Asulina" per km., while the gasoline consumption in parallel tests was 1.1 per km. The price of "Asulina" is half that of gasoline. W. L. OWEN.

Suitability of fuels for Diesel engines. LOUIS R. FORT. *Power* 73, 831-5 (1931).—Since the factors which control the behavior of fuel in the engine cylinder are not well understood, performance tests must be relied upon in selecting suitable fuels. D. R. DINA.

Hydrogenation at high pressures. GEORGE ROBERTI. *Mém. Acad. Italia (Classe Sci. Fis. e Mat.)* 1, Chpt. No. 2, 15 pp. (1930).—As part of the general problem of utilizing national sources of fuels, the hydrogenation of the phenolic distillate from lignite was studied. Co_2S_3 was found to be a satisfactory catalyst, as it induces the formation of said cyclohexane at 100 atm. at temps. of $350-400^\circ$, and, at the same time, is much less sensitive to poisoning than other commonly employed catalysts. A. W. C.

The possibilities of a South African oil industry. O. FULMANN. *S. African Mining and Eng. J.* 41, 17-11, 251-2 (1930).—F. discusses the low-temp. carbonization of S. African lignite followed by hydrogenation of the tar produced. A. H. B.

The history of gas. C. I. VAN BUREN. *Ind. Gas* 51, 40-6 (1931). R. J. C. VAN DER HORVEN.

Natural gas problems. ALEXANDER FORWARD. *Gas Age-Record* 66, 823-6 (1930).—A discussion. LESLIE H. BRAGO.

Statistics of the British manufactured gas industry. DAVID BROWNIE. *Gas Age-Record* 66, 975-7 (1930).—A review. LESLIE H. BRAGO.

The Lins gas works—its development and its future position. LUDWIG STOLL. *Z. Berg. u. Hüttenw.* 74, 70-40 (1931). R. H.

Nemography in gas analysis. W. J. G. DAVEY. *Gas World* 92, 604-5 (1930).—The graphical method is used for the detn. of H_2 and CH_4 in a gas after the other constituents and the caloric value have been detd. The other constituents usually detd. are CO_2 , O_2 , unsatd. hydrocarbons, CO and N_2 . Examples are given showing the application of the method to carbonated water gas and vertical retort gas. F. H. B.

Institution gas research fellowship report 1929-30; Influence of the inorganic constituents in the carbonization and gasification of coaly liberation of ammonias. A. H. EASTWOOD AND JOHN W. COHN. *Inst. Gas Eng., Communication No. 23*, 24 pp. (1930).—See C. A. 25, 1037. G. G.

Heat transmission: modern methods of expressing convection data. M. FRUENKEL. *Gas World* 92, 602-7 (1930).—The principles of heat transfer are reviewed and expressed in math. form. Many heat transfer problems connected with the gas and carbonization industries are too complex to be expressed in simple generalized rules. In some individual cases it is possible to correlate and apply available data. The math. presentation is supplemented with graphs for the solution of heat-transfer problems, particularly those in which convection is taken into account. F. H. BIERGREN.

Gas meter calibration with the Mariotte bottle. L. ZIRKER. *Gas Wessersf.* 74, 27-8 (1931).—An insulated Mariotte bottle is a convenient means for calibrating

gas meters. An illustration of such use and the method of calculating exact gas vol. are given with examples. R. W. RYAN

The new (large) Brandl-Marischke gas meter. PITTSBURGH EQUITABLE METER CO. *Gas u. Wasserfach* 74, 67 (1931), cf. C. A. 24, 5435.—The principle of the Brandl-Marischke meter is old and has long been used by the Pittsburgh Equitable Meter Co. Reply BRANDL AND MARISCHKE *Ibid* 67-70.—Patents granted in various countries are listed. R. W. RYAN

Coal-grinding and -mixing equipment for gas works. OTTO HUPPERT. *Gas u. Wasserfach* 74, 44-56 (1931).—When gas works are large enough there is a definite advantage in grinding and blending coal, as cheaper coal and waste coal and coke can be used and a good coke still obtained. Operating and test data are given in support of this claim and a typical German installation is illustrated and discussed. Attempts to use low temp. brown-coal coke in such blends were unsuccessful. Recent work of Damm, Fischer and others on the mechanism of coking is reviewed. R. W. RYAN

The physical basis and the manner of operation of the Progas-Union gas governor. C. BÄCKEMANN. *Gas u. Wasserfach* 74, 110-2 (1931), cf. C. A. 25, 100.—Addns. and correction to previous math. derivations are given. R. W. RYAN

Experiences with a waterless gas holder. FRANK PRENTICE. *Inst. Gas Eng. Communication No. 7*, 6 pp (1930).—See C. A. 24, 4918. G. G.

Gas dehydration. E. W. SMITH. *Inst. Gas Eng. Communication No. 5*, 37 pp (1930).—See C. A. 24, 4917. G. G.

Effect of sodium carbonate upon gasification of carbon and production of producer gas. DAVID FOX AND ALFRED H. WHITE. *Ind. Eng. Chem.* 23, 239-66 (1931).—Na₂CO₃ reacts in appreciable measure with C at temps. above 800° and the rate becomes rapid above 900°. The reactions are Na₂CO₃ + 2C = 2CO + 2Na, Na + CO₂ = NaO + CO, NaO + CO₂ = Na₂CO₃. The net effect is to increase the vapor pressure of C to 10¹⁴ times its actual value. The mechanism of the reaction is discussed. The use of coke impregnated with Na₂CO₃ in the gas producer gives gases higher in CO than would otherwise be obtained. In a water gas machine the gases are much richer in CO and H₂, but the increased reactivity of the fuel bed would be disadvantageous during blasting and off-set any gains during steaming. ALFRED H. WHITE

The new Omeco fine-gas tester. JULIUS FEUERUNGSTECH. 18, 242 (1930).—This recorder tests O₂ instead of CO₂. An automatic device takes in predet. vol. of the sample and of H₂ and passes them through a combustion furnace. The contraction is recorded. ERNEST W. TIMELX

A new method of phenol recovery from gas liquors. C. SCHÖNBURG. *Brennstoff-Chem.* 12, 69-71 (1931), cf. *Glückauf* 64, 436 (1928).—The disadvantages of dephenolizing with C₁₂H₆ are discussed. Tricresyl phosphate has been found to absorb 10-20 times more phenols and also has the advantages of non-volatility, invol. in, and greater sp. gr. than the liquors. Partition coeffs. for both are shown diagrammatically for phenol and cresol solns. and gas liquor. The phenols, after absorption, are driven off from the high boiling tricresyl phosphate by heat or steam distn. under vacuum. Practice shows that only 4-5 parts of tricresyl phosphate are needed to dephenolize 100 parts of liquor contg. 3 g. phenol per l. and that a 6-8% phenol concn. so obtained in the phosphate requires only 5-6 parts steam per part phenol for complete recovery. Use for several months shows a selectivity for phenols and the non soln. of tar. App. shown schematically, is described. F. W. JUNG

Tests on a Birmingham (town gas-fired oven) furnace installation. C. M. WALLER. *Gas J.* 190, 575-7 (1930).—Tests were carried out on a battery of 4 town gas fired oven furnaces to det. how far working efficiencies could be affected by the employment of special insulating material, improved recuperator elements, and in particular the precise control of primary and secondary air, and reduction of losses by leakage through doors and dampers. For a given operation of carbonizing in the older type of furnace a fuel consumption of 16,000 cu. ft. of 500 B. t. u. gas was required. The improved installation used 6,230 cu. ft. of gas of slightly lower calorific value. F. H. H.

A new commercial application of the Walter Feld polythionate process. FRITZ OVERACKER. *Z. anorg. Chem.* 43, 1048-51 (1930).—To remove NH₃ and S from coal gas the following reactions have been used (NH₄)₂S₂O₈ + 2NH₃ + H₂S → 2(NH₄)₂S₂O₅ + S, (NH₄)₂S₂O₅ + 2NH₃ + H₂O → (NH₄)₂S₂O₄ + 2NH₃ + H₂O, (NH₄)₂S₂O₄ + S, (NH₄)₂S₂O₄ + 3H₂S → (NH₄)₂S₂O₃ + 3S + 3H₂O, (NH₄)₂S₂O₃ (not) → (NH₄)₂SO₄ + SO₂ + 2S, or (NH₄)₂S₂O₃ + 2(NH₄)₂SO₄ → 3(NH₄)₂SO₄ + 2S. The S is burned to SO₂ for re-use. FUNK (Gas u. Wasserfach 68, 383 (1925)) and Raschig (C. A. 15, 440) have investigated the process. A modification at Leverkusen makes possible complete

removal of NH_3 and H_2S without addn of foreign NH_3 ; if excess H_2S is present, by gas washing in several stages. H_2S alone reacts very slowly with $(\text{NH}_4)_2\text{S}_2\text{O}_8$, but rapidly when NH_3 is also present. Absorption of excess H_2S (without NH_3) is done in the first stage, where H_2S is coned, while in later stages the more difficult absorption of dil H_2S is accomplished by the aid of NH_3 by the first reaction given. E. M. SYMMES.

A new modification of the Feld polythionate process. D. STAVORINUS. *Helv. Chim. Acta* 14, 19-20 (1931).—A short review of a new modification of the Feld process as developed by the I. G. and used at the mine Auguste Viktoria (cf. Overdick, preceding abstract). It consists of cooling the crude gas from the ovens to 25° , thus removing tar and part of the NH_3 , and then removing both H_2S and NH_3 in 4 scrubbers by polythionate liquor circulating in the same direction as the gas. In the first scrubber H_2S is taken out to an extent of more than the equiv. of NH_3 present; in the second scrubber no NH_3 is present, but thionate and H_2S react to give free S. Furthermore, the regenerated liquor (treated with SO_2) contg. some free SO_2 is circulated over this scrubber. In the third scrubber is used a mixt. of thionate liquor with the NH_3 distd. off the primary cooler condensate; the gas is thereby freed of its last traces of H_2S . A final washing for removal of the remaining NH_3 is carried out in the fourth scrubber run on NH_3 free thionate liquor, which is subsequently used in the first scrubber. The spent liquor is worked up in the customary way. S is filtered off and the liquor decomposed at $120-30^\circ$ to give a strong $(\text{NH}_4)_2\text{SO}_4$ soln., which is freed of thiocyanate by fractional crystn. The present process is an improvement over the former ones in that H_2S comes in contact with liquors of high NH_3 content. This has been found necessary for complete H_2S absorption. The thio-sulfate formed from the tetrathionate of the active soln. during the reaction with NH_3 and H_2S is regenerated by the action of SO_2 -contg. gases in a separate little scrubber. D. J. C. VAN DER HOEVEN.

The storage of carbon monoxide. OTTO HEINRICH WAGNER. *Brennstoff-Chem.* 12, 87-9 (1931).—In view of the possible use of CO as a motor fuel, three methods of storage have been investigated using (1) liquid absorbent, (2) solid absorbent and (3) a reversible chem. combination. Ether-pentane mixts. were the best of the liquids (absorption increasing with decreasing b. p.), but liquid CO_2 and NH_3 showed no absorptive properties. Of the solids, activated C when packed in cylinders showed only 40% capacity increase over an unpacked cylinder at 40 atm., and less at higher pressures. Silica gel and chabasite showed a decrease in capacity at pressures under 100 atm. Solns. of ammoniacal and acid cuprous salts involved heating to deliver the absorbed CO and were corrosive. With solid cuprous halides, addn. compds. of the type $\text{Cu}_2\text{X}_2 \cdot 2\text{CO}$ were formed very slowly during the first absorption but rapidly after a single removal of the absorbed CO. The addn. is quant. reversible and traces of H_2O greatly increase its speed. Purification as well as storage is thus feasible. The vapor pressures of the COCl_2 and COBr_2 showed 235 mm. Hg for the former at 20° and 1.67 atm. for the latter. A 40-l. cylinder holding 7 kg. CO at 150 atm. will contain 28 kg. when filled with 100 kg. solid Cu_2Cl_2 . Warming to 50° then releases the CO for use, the vapor pressure of the addn. compd. being 5 atm. at this temp. Because of the cost of Cu_2Cl_2 , the halides and ammoniates of Fe, Ni and Co as well as ZnO and CaO are being studied. F. W. J.

Ammonia synthesis by the low-pressure Mont-Cenis process. W. F. SCHOLVEN. *Chem. Met. Eng.* 38, 82-8 (1931).—The Scholven, Germany, plant is described. Coke-oven gas is the source of H. CO is removed from the gas by washing with liquid N. Conversion is effected at 400° and 100 atm. Economic data and advantages of the process are given. L. W. T. CUMMINGS.

Development and inventions in tar distillation. ERNST WEISS. *Asphalt u. Teer* 30, 1124-7, 1147-9, 1169-70, 1232-5, 1287-92, 1312-5 (1930). cf. C. A. 24, 3886.—This is a detailed digest of 106 recent patents. K. H. ENGEL.

Specifications for coal tar in Holland. F. J. NELLENSTREIJN. *Asphalt u. Teer* 30, 982-3 (1930). K. H. ENGEL.

Analysis of coal-tar-asphalt mixtures. H. MALLISON. *Asphalt u. Teer* 30, 183-4 (1930).—Sepn. of the components by sulfonation methods is not entirely trustworthy. The asphalt is converted into sol and insol sulfonic acids. Certain asphalts, sulfonated by themselves, showed slight increases in wt. When sulfonated in mixts. with 85% tar, they were changed, almost quantitatively, into sol products. K. H. ENGEL.

Thermal decomposition of coal-tar constituents. VII. Reaction products of the thermal decomposition of m-cresol. YOSABURO KOSAKA. *J. Soc. Chem. Ind., Japan* 34, Suppl. binding No. 1, 10-2 (1931), cf. C. A. 24, 1938.—Pure m-cresol was decomposed at 700°, 800° and 900°, as was done with C_{11}H_8 (*J. Soc. Chem. Ind., Japan* 31A, 507, 31B, 126 (1928)) except that at 900° no packing material was used because of C deposits formed. Fractions of the distillate up to 120° were mainly C_{11}H_8 and PbMe. Fractions

above 120° were treated with NaOH soln to sep acid constituents, the remainder was sepd. into picrate forming and non picrate forming substances. The former were mostly naphthalene, anthracene, phenanthrene and derivs, the latter mostly Ph₂, Ph₂O and derivs. Differences in packing material have little effect. The yield of C₁₀H₈, PhMe and PhOH is a max. at about 800°, the yield of naphthalene and anthracene varies little with temp., but the Ph₂ yield increases with decompn temp. VIII. React. on products of thermal decomposition of o-cresol and p-cresol. Ibid 12-3 — Pure o-cresol and p-cresol were decompd. at 800° and the products compared with those from m-cresol treated similarly. The main products were C₁₀H₈, PhMe, PhOH, naphthalene, anthracene, phenanthrene, Ph₂, Ph₂O and their derivs. Nearly 21% of m-cresol, 12% of o-cresol and 8% of p-cresol remain undecompd. The PhOH yield is 19% with o-cresol and p-cresol, but 4.6% with m-cresol. m-Cresol gives the most C₁₀H₈ and PhMe while o-cresol gives the most naphthalene and anthracene. IX. Qualitative consideration of the reaction mechanism of the thermal decomposition of cresols. Ibid 13-4 — The thermal activity of cresols has been considered previously (Fischer, *Abh d Kohle* 6, 172(1923), Hagemann, *Braunkohle* 28, 1095(1929)). A series of reactions based on exptl. data is proposed. The difference in nature of decompn. of the 3 cresols seems to be due to the relative stability of OH and CH₃ groups, CH₃ being more stable than OH in m-cresol, and vice versa in o- and p-cresol. E. M. SYMMES

Light oil from water-gas tar. N. A. ORLOV. *Zhur. Prikladnoi Khim.* 3, 535-7 (1930) — Light oil obtained in the manus. of water gas which had a brownish color and d₄²⁰ 0.9249 was distd. and the 141-50° fraction was taken to sep styrene. This fraction, which had d₄²⁰ 0.8704 and n_D²⁰ 1.5121, on being treated with Br in chloroform soln. indicated the presence of about 30% of styrene. The bromination was carried out at 0° and the remaining hydrocarbons were distd. off with steam, the dibromide was twice recrystd. from alc. and m. 71-1.5°. In a vacuum distn. at 25 mm. pressure the whole of the dibromide distd. over at 150-2° and required only one recrystn. The light oil contains about 1.5-2% styrene. This method, however, is not recommended because other products present are also brominated. The pure styrene was obtained from the dibromide by the action of Mg. Attempts to obtain styrene without the use of Br by heating the fraction in an autoclave for 5 hrs. at 200° failed. Indene was easily sepd. from the 176-85° fraction by the Weingerter method (cf. C. A. 3, 1152) with a yield of 42 or 47% corresponding to that from the original light oil. The indene content is probably higher because the liquid becomes very hot during the reaction and polymerization can be noticed. A. A. BOEHLING

Tar dehydration. EMLERT. *Gas u. Wasserfach* 74, 40(1931) — The longest possible tar well is desirable. Heating coils were so placed as to minimize mixing by convection. A week before tar was to be loaded it was heated in about 24 hrs. to 50° and then allowed to stand without further heating for 3 days before loading. R. W. RYAN

Functions of coke ovens. EDWARD G. STEWART. *Inst. Gas Eng., Communication* No. 6, 20 pp (1930) — See C. A. 24, 5139. G. G.

Production, properties and use of coke. D. AUFLANDER. *Montan. Pundschau* 22, 521-4(1930) — Some statistics are given on the production of gas, coke and coal tar products in Germany. For gas making the C content of the coke should be high and the H content low. The size of lumps recommended is 2-9 cm. diam. (0.75-3.5 in.), depending on the size retort used. CHANNING WILSON

Present position of the metallurgical coke industry in Central Europe. CH. BERTHELOT. *Rev metal* 27, 496-500(1930) — The present tendency in the Central European metallurgical coke industry is to improve furnace design and furnace equipment, and to analyze more intensively factors entering into the net cost of coke manus. C. H. LORIO

The reactivity of metallurgical coke. DUFRANKE. *Rev metal* 27, 509-11(1930); cf. C. A. 24, 1930 — The surface of coke was modified by applying solns. of CaO and flux. The reactivity of 3 cokes to which 3 solns. were applied was detd. for 800°, 900°, 1000° and 1100°. Lime solns. decrease the reactivity of coke, thereby raising the CO₂ content of the combustion gases. C. H. LORIO

The determination of coke strength. FRITZ G. HOFFMANN. *Brennstoff-Chem* 12, 61-5(1931), cf. C. A. 24, 1931 — The effects of fissuring, cooling or quenching strain and coke hardness upon the accuracy of the results obtained in tumbler, compression and shatter tests are discussed in detail. Causes of fissures in coke from various types of ovens, as well as in the carbonizing of briquetted coal, are studied. The use of an automatic screening, weighing and recording device is suggested for continuous detns. on a small part of the oven output. The tumbler with attached double screen is recom-

mended as the best-suited hardness test. The importance of testing carefully screened coke and the preferability of hand- to fork picked samples is stressed. F. W. JUNG

Low-temperature pressed coke. E. ROSER. *Brennstofftech. Chem.* 12, 88-7(1931).—By using thin-walled, slightly tapered alloy steel cylinders of 300 mm. mid-diameter and 2800 mm. height, fitting in pairs into 400-mm. Otto ovens, a series of carbonizations at oven temps. of 750° has been made. A strongly swelling slack of 18% volatile carbonaceous matter (V. C. M.) and 5.4% ash was coked 4 hrs. A coking slack of 22.1% V. C. M. and 5.1% ash, and a 1:1 mixt. by vol. of this slack and an anthracite fine of 10.5% V. C. M. and 7% ash were coked 5 hrs. In each case a hard, large-sized and uniform coke resulted. Gas and tar yields were not detd. Analytical data are given for the coke obtained. Consumption and production data for light oil products in Germany are given in support of a low-temperature process. F. W. JUNG

Dry quenching of coke. WM. O. RENTON. *Fuels and Steam Power* (A. E. S. M. Trans.) 53, 65-85(1931)—See C. A. 25, 1039 G. G.

The reductivity of over-crusted coke (PIKOWARSKY) 9. Ammonia liquor as a fertilizer (BIRCHALL) 15. The microbiological aspect of peat formation (THAYSEN) 8. The present nature of coals resulting from their history (DUBREIL) 8. The dehydrating action of coal ash (TANAKA) 18. Separation and synthetic preparation of phenol (MIROSHITA) 10. Apparatus for eliminating CO from motor-vehicle exhaust gases (U. S. 1,791,912) 1. Separation of tars, tar oils, etc. (Brit. pat. 336,945) 22. Purifying light oils recovered from coal gas, etc. (Brit. pat. 336,636) 22. Clarifying plant for coal industry (Ger. pat. 516,454) 1. H and CO from CH₄, etc. (Brit. pat. 336,635) 18. Separation of dust from gases (Brit. pat. 336,201) 13. Purifying light hydrocarbons produced by destructive hydrogenation of coal (Brit. pat. 335,885) 22. Purification of gases, etc. (Belg. pat. 371,622) 4.

Fuel. WILHELM CLEMENS. Ger. 513,512, July 1, 1926. Dried pressed peat is pulverized, kneaded into a plastic condition with added substances and finally pressed into molds. The pulverized mass may be coked. App. is described.

Fuel briquets. SOC. ANON. LA CARBONITE. Ger. 513,731, May 6, 1924. Briquets of high caloric value, suitable for driving gas producers, etc., are obtained by distg. wood, peat or lignite below 500°, mixing the residue with oxidized tar below 500° and pressing. Cf. C. A. 24, 1940

Fuel for motors. N. V. DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ. Fr. 604,033, Apr. 16, 1930. Fuel suitable for motors is obtained by heating carbonaceous materials such as petroleum, asphalt, etc., in the presence of a compd. of Mo acting as a catalyst and in the presence of H under pressure, the conditions being chosen so that no splitting of the mols. of the treated materials takes place with formation of gaseous or low b. p. products. Examples are given.

Fuel for internal-combustion engines. JOHANN W. MICHEL. Ger. 515,816, Nov. 20, 1929. Addn. to 515,076 (C. A. 25, 1660) Soot is mixed with powd. fuels giving a relatively large amt. of ash, e. g., lignite, wood charcoal or cellulose.

Vaporizing heavy fuels for internal-combustion engines. RECTOR GASIFIER CO. Brit. 336,243, July 9, 1929. The fuel is heated, sprayed radially to a common center from several orifices and an inert gas is passed at right angles through the sheet of spray thus formed.

Catalytic heating device for fuel of internal-combustion engines. OSCAR H. WALLIN. U. S. 1,792,337, Feb. 10. Various structural details are described.

Removing carbon deposits from internal-combustion engine cylinders. THOMAS & HOCHWALT LABORATORIES, INC. Brit. 335,943, June 5, 1929. A C-loosening compn. is used contg. a halogenated aromatic hydrocarbon (such as chloronaphthalenes, chlorobenzenes, bromobenzenes, halogenated toluenes or xylenes, benzyl chloride or benzyl chloride, which may be dissolved in alcs., ether, acetone, EtOAc, kerosene, gasoline, C₆H₆ or the like) which is introduced into the hot cylinders. Examples with proportions are given.

Acetylene. CHARLES DE LA ROCHEFFE. Ger. 513,692, Mar. 6, 1929. Details are given of a plant for producing an C₂H₂-air mixt. for motors and other combustion machines.

Solid acetylene. I. G. FARBENIND. A.-G. Fr. 604,001, April 15, 1930. C₂H₂ is stored and transported by cooling the gas to just below its m. p. and compressing the snow thus formed into cakes.

Wagon for the low-temperature distillation of bituminous fuels in tunnel ovens. N. V. MASCHINERIEËN EN APPARATEN FABRIEKEN. Ger. 516,153, Oct. 18, 1927.

Destructive hydrogenation. IMPERIAL CHEMICAL INDUSTRIES, LTD., R. HOLROYD and C. COCKRAM. Brit. 337,028, July 18, 1929. At the beginning of a multi-stage process for the destructive hydrogenation of bituminous coal, the coal is heated either *in vacuo* or in an inert gas or with H₂ with or without oil or catalysts, to a temp. at or just above that of initial thermal decomposn. at which the plimic constituents disintegrate. Preferably H₂ is used under pressure in the presence of high-temp. tar or oils derived from a later stage of the process and in the presence of catalysts, and sufficient H₂ is used to deoxygenate the coal and yield a fusible product. The material is then further hydrogenated and liquid products are fractionated from the resulting substances. Various details and examples are given.

Destructive hydrogenation. C. P. R. HARRISON and IMPERIAL CHEMICAL INDUSTRIES, LTD. Brit. 336,901, July 19, 1929. Destructive hydrogenation under pressure is effected in a vertical tower in the presence of a massive metallic catalyst of large surface area such as plates, wire, ribbon or gauze. Various details of app. are described. Cf. C. A. 25, 192.

Destructive hydrogenation of pastes of coal and oil. C. F. R. HARRISON, E. D. KAMM and IMPERIAL CHEMICAL INDUSTRIES, LTD. Brit. 336,610, July 16, 1929. Various details of app. and procedure are described for a process in which all the products from the reaction vessel pass through a common exit to a catch pot which is kept at such a temp. (suitably about 350°) that heavy oils are condensed while the vapors of middle and light oils are led away and the heavy oils are returned to the reaction vessel.

Furnace and retort for drying, roasting or distilling solid materials such as coal. F. K. WOODROFFE. Brit. 336,657, July 18, 1929.

Coal. MAURICE DÉLOS. Fr. 694,979, May 5, 1930. The combustion properties of coal are improved by sprinkling it with water contg. soda, S, red ochre and CaCl₂.

Apparatus for cleaning coal by air separation. ALWYN C. OTTY and GRANGER B. SADLER. U. S. 1,792,440, Feb. 10. Structural features.

Apparatus for cleaning and purifying coal. HARRY L. McLEAN (one third each to George W. Wilmet and Francis H. Blatch). U. S. 1,792,179, Feb. 10. Structural features.

Washer for coal, ores, etc. ANTOINE FRANCE. Ger. 516,145, Nov. 21, 1926.

Liquid products from coal, etc. N. V. DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ. Fr. 694,192, Apr. 18, 1930. Valuable liquid products are prepd. by treating coal, tar, petroleum products or residues, lignite or cellulose by heating under pressure in the presence of H₂ and a catalyst composed of Mo compds. intimately mixed with finely divided absorbent substances such as absorptive charcoal.

Apparatus for determining the humidity of brown coal, etc. HANS TRAUTHWEIN. Ger. 513,701, April 6, 1928. The water is evapd. in a closed chamber in a rotary drum app. and the condensate is measured.

Lignite cooling plant. MASCHINENFABRIK BUCKAU R. WOLFF & CO. Ger. 513,860, Jan. 26, 1929. Addn. to 460,570. Cooling plates are arranged after the manner of venetian blinds.

Drying peat. WILHELM CLEMENS. Ger. 513,572, Jan. 25, 1927. To crude peat pressed peat is added, and the whole is further pressed and then treated with drying agents.

Montan wax. I. G. FARBENFABRIK A. G. Brit. 335,929, July 1, 1929. Purification of montan wax is effected by extrn. under pressure and at an elevated temp. (suitably 135° or 150°) with org. solvents of the wax acids and esters (but not of the coloring substances) such as aliphatic alcs, esters, ketones and ethers, open chain or cyclo-aliphatic liquid hydrocarbons or chlorinated hydrocarbons, alone or mixed or with a smaller quantity of aromatic hydrocarbons. Examples with details of procedure are given.

Removing phenols from ammonia liquor. GEWERESCHAFT MATTHIAS STENNES. Ger. 517,001, Apr. 30, 1929. In removing phenols from ammonia liquor by means of a circulating org. solvent, the latter is desulfurized, before freeing it from phenols, by treatment with a suitable liquid or gas. Thus, the solvent may be washed with water, which may then be passed to the ammonia washer, or aerated to render it fit for use again. Alternatively, the solvent may be treated with coke-oven gases, which may traverse a cycle in which the H₂S taken up by the gases is removed by a solid purifying material or by an alk. wash. Cf. C. A. 25, 801.

Gas purification. P. BÖNNEMANN. Brit. 336,425, Oct. 12, 1929. See Fr. 682,838 (C. A. 24, 4616).

Gas purification. DAVID L. JACOBSON (to Koppers Co.). U. S. 1,792,097, Feb. 10. Gas is washed with an alk. suspension of a compd. of a metal of the Fe sub-group of the 8th group of the periodic system, such as Fe oxide, contg. in soln. a hydroxy deriv. of

an org. non-paraffin compd., such as 2% or less of a phenol, which retards formation of thio compds. *Cl. C A 25, 802*

Purifying gases. I. G. FARBENIND. A-G. Brit. 336,310, July 20, 1929. Coke-oven gases, low temp. distn. gases, cracking gases and gases other than those mentioned in Brit. 329,698 (*C A 24, 6943*) such as gases intended for the manuf. of synthetic NH_3 , MeOH , etc., are freed from volatile S compds. such as H_2S and largely from CO , if present in appreciable proportion, by washing with high b. p. hydrocarbons or phenols or their mixts. under 10 atm. or higher pressure at such a temp. that the concn. of H_2S amounts to less than 0.3% by volume of the gas. The pressure is preferably released in stages. Various details and examples are given.

Purifying gases. HEINRICH KOPPEL A-G and CHRISTIAN J. HANSEN, Ger. 513,913 Nov. 1, 1927. NH_3 and H_2S are removed from gases in the proportion of 2:1 by a washing liquid causing the formation of Fe, Mn or Zn thionate, and the excess of H_2S is removed by a second washing causing the formation of polythionates. H_2SO_3 is added to the second washing fluid. *Cl. C A 24, 1726*

Gases rich in olefins. I. G. FARBENIND A-G (Fritz Winkler, Paul Feiler and Carl Meserveknecht, inventors). Ger. 513,815, Nov. 11, 1928. Gases rich in olefins are obtained by rapidly heating powd. bituminous coal to dull red heat and rapidly cooling the gas evolved. Superheated steam may be admitted to the glowing fuel. The examples describe the production of gas contg. 10 to 21% C_2H_4 from lignite by this method.

Heating and illuminating gas. ALBERT F. KUNDEAGER (to United Gas Improvement Co.) U. S. 1,792,172, Feb. 19. A mass of coal is heated to incandescence with an admixt. of finely divided iron or Fe oxide in a chamber, and a mixt. of H and superheated steam is introduced into the heated mass, simultaneously addnl. coal is introduced into the chamber together with an admixt. of a small proportion of finely divided Fe oxide, and the gases within the chamber are subjected for a time to steam pressure of 30-300 lbs. per sq. in. to cause the H to combine with the C to form hydrocarbon gases. App. is described.

Coal gas. GÖR. LINNÉ'S EISMASCHINEN A-G (Paul Schulten, inventor). Ger. 513,815, Aug. 7, 1928. The small quantities of reactive gas remaining in coke-oven gas are removed, before splitting the gas by cooling to low temps., by passing it over an absorptive agent such as active C impregnated with metal or metallic salts, or over active SiO_2 -gel. This removes the NO and traces of hydrocarbons remaining after the removal of the hydrocarbons, H_2S and water by the usual methods.

Washing coal gas. H. KOPPEL A-G. Brit. 336,481, Sept. 27, 1929. In final water cooling of coal distn. gases, the water, before re cooling, is washed or treated with tar to remove deposited naphthalene. App. is described.

Expansion machines for separating benzene from coal gas. GÖR. LINNÉ'S EISMASCHINEN A-G (Paul Schulten, inventor). Ger. 513,764, Sept. 14, 1928. Addn. to 511,567 (*C A 25, 1205*). Details are given for lubricating the machine by oil to which homologs of C_8H_{18} or hydrated naphthalene are added to lower the congealing point. Glycol, ketones or monohydric alcs. may be added to the water-ice soln. to prevent excessive loss by evapn. at low-vapor pressures.

Mixed oil gas and water gas. J. A. PERRY (to Humphreys & Glasgow, Ltd.) Brit. 336,525, March 2, 1929. In the operation of a carbureted water-gas set, oil gas is obtained by passing sep. streams of a hydrocarbon fluid through the generator and through the carbureter, the hydrocarbon material being cracked by passage through the fuel bed, and the gas is withdrawn separately from that formed in the carbureter and superheater. Various details of app. and operation are described. *Cl. C A, 25, 1003, 1662*

Water gas. METALLGES. A-G and C. MUELLER. Brit. 336,821, Nov. 4, 1929. Powd. and granular fuels are gasified by O in admixt. with steam, and part of the gas produced is reintroduced into the producer together with the gasifying medium, in which it is burned with a part of the O before the gasifying medium reacts with the fuel to be gasified.

Non-poisonous gas from blue water gas. W. H. FULWEILER (to Humphreys & Glasgow, Ltd.) Brit. 335,869, May 2, 1929. One portion of blue water gas is passed over a heated catalyst such as Ni and another portion is passed with steam over a heated catalyst such as activated Fe oxide to form chiefly H_2 , the 2 gases are mixed, with removal of CO_2 and water vapor either before or after admixture. Some CO may remain.

Continuous water-gas production. J. PINTSCH A-G. Brit. 336,965, July 22, 1929. In the continuous production of water gas by passing steam and water gas together with the fuel through a tube surrounded by a heating app. for the reacting

gases (as described in Brit. 205,717 (C. A. 23, 2278)), the mixt. of previously produced water gas and steam is admitted progressively and cumulatively to the reaction tube at a number of points along the tube as well as through a space at the top of the tube. Details of app. are described.

Water-gas generator. W. B. CHAPMAN, C. W. ANDREWS and H. B. YOUNG (to H. A. Brassett & Co. and Western Gas Construction Co.) Brit. 335,203, Nov. 27, 1928. A generator is described having means for continuous ash removal, in the operation of which the air steam ratio is such as to maintain a zone in which the ash is liquefied and a cooler zone below this in which the ash tends to solidify for removal by a mech. device.

Water-gas generator. H. G. TERZIAN (to Humphreys & Glasgow, Ltd.). Brit. 336,902, April 20, 1929. A generator using bituminous fuel has a perforated tubular gas off take in the axis of the upper part of the generator and openings communicating with an annular chamber surrounding the carbonizing zone, and both blast and run gases pass through this annular chamber. Various auxiliary structural details also are described.

Combined water-gas and steam generator. CARL MARISCHEK. U. S. 1,702,178, Feb. 10, 1928. Structural features of an app. with a boiler chamber surrounding a generator shaft, a heating jacket surrounding the boiler chamber, and tubes disposed entirely within the heating chamber and connected at their ends with the boiler chamber.

Gas producer. J. A. MACDONALD. Brit. 330,932, July 20, 1929.

Gas producers. PORTER G. M. & H. Fr. 694,118, April 17, 1930. See Ger. 503,750 (C. A. 24, 5095).

Vertical gas producer. C. H. LANDER and T. F. HURLEY. Brit. 336,022, July 10, 1929.

Grateless gas-producer for the continuous gasification of coal dust or granules. JANOS KALICZEY. Austrian 120,427, July 15, 1930.

Gas producer and associated furnace and boiler. BARCOCK & WILCOX, LTD. Brit. 336,356, Aug. 22, 1929.

Charging apparatus for gas producers. A. L. GALUSHA. Brit. 337,005, July 23, 1929. Structural features.

Degasifying chamber for gas producer. CARL BLEYER. Ger. 610,351, Feb. 8, 1927.

Prevention of gummy deposits in gas mains, etc. W. H. FULWELER (to Humphrey & Glasgow, Ltd.). Brit. 330,456, Dec. 17, 1928. Gum forming constituents such as indene and styrene are left in the gas and the latter is treated with mono- or poly-hydric derivatives of $C_{11}H_8$ or its homologs, as by spraying.

Tar distillation. GEORGE T. GAMMILL, JR. (to The Barrett Co.). Can. 309,014, Mar. 3, 1931. A method and an app. for distg. tar are specified. The tar is heated without sepn. of the vapor, then the vapor is released and simultaneously the tar is blown with an inert gas preheated to a temp. approx. that of the tar, the vapor and inert gas are withdrawn, the mixt. is cooled to sepn. condensable constituents, and the inert gas is heated and returned for further contact with addnl. portions of the tar. Cf. C. A. 25, 803.

Coking coal. A. POTT. Brit. 335,910, April 27, 1929. A coking coal is first dried and heated externally to a temp. just below its decomposition temp. and is then transferred to another app. in which it is externally heated to effect distn. at a temp. not exceeding 550° (the second app. being preheated to about the temp. of the coal before introduction of the latter into it).

Wet or dry cooling plant for coke. C. OTTO & Co. G. M. & H. Ger. 516,442, Aug. 3, 1926.

Coke ovens. RUDOLF WILHELM. Fr. 694,019, April 16, 1930. Construction of door is described.

Coke ovens with heat regeneration. ROBERT J. A. DE STORBEUR. Fr. 693,954, Dec. 14, 1929.

Coke oven for making semi-coke. SOC. GÉNÉRALE DE FOURN. À COKE SYSTÈMES LECOQ. Brit. 336,809, Oct. 27, 1928. Structural features of app. with a vertical or inclined chamber oven lined with metal plates or partitions supported at a distance from the heated refractory walls and heated by radiation only (each metal chamber being enclosed on all its faces by the heating walls).

Multiple regenerative coke oven. C. OTTO & Co. G. M. & H. Ger. 513,593, May 30, 1928. Details of arrangement.

Coke-oven construction and operation. K. BEUTHNER. Brit. 337,278, Jan. 3, 1929. Distillates are withdrawn from the bottom of the charge into a lower compartment forming part of a divided fuel support, whence they pass to an outlet pipe. Steam may be introduced for production of water gas. Cf. C. A. 24, 2868.

22—PETROLEUM, LUBRICANTS, ASPHALT AND WOOD PRODUCTS

W F PARAGHER

Artificial aging of mineral oils. III. F. EVERS AND R. SCHMIDT *Wiss. Veröffentl. Siemens Konzern* 9, No. 1, 357-68 (1930), cf. *C. A.* 23, 3565, 25, 404.—The catalyst and procedure previously recommended for aging tests on mineral oils are used to show that the aging curve may be expressed by the equation $y = 2.3 \log x \cdot C + K$, where $C = 10.1$ and $K = -11.3$. From the aging const. of a mineral oil, the sapon value, acid value and tendency to polymerization during various stages in aging can be calculated, and a close approximation made of the av. life of the oil under elec. conditions. B. C. A.

Heat calculation for flash distillation [of petroleum hydrocarbons]. G. H. FANCHER *Chem. Met. Eng.* 37, 307-8 (1930).—In calcg. the heat added during vaporization of hydrocarbons by means of Turner and Harrell's equation (*C. A.* 24, 2280), more reliable results can be obtained by taking into consideration the av. sp. gr. and the temp. rather than the mol. wt. of the vapor formed. The value of the latent heat used in this calcn. should be obtained from the relation $l = (1/d)(110.9 - 0.09 t)$, where l is the latent heat of vaporization in B. t. u. per lb., d the sp. gr. (at 60°F. compared to water at 60°F.) of the liquid vaporized and t is the temp. in °F.; this value is more accurate than that obtained from Hildebrand's or Trouton's rules. B. C. A.

Determination of carbon residue in petroleum products—Conradson carbon. B. ROSENBAUM *Chem.-Ztg.* 55, 106 (1931).—Corrections are given for several Conradson C detns. previously reported (*C. A.* 25, 804), with the explanation that the quality of these products has changed in the meantime. The possibility of a definite relation between Conradson C and viscosity is pointed out. D. F. BROWN

Ultrafiltration of petroleum. A. ZAHARIA AND E. LUCATU. *Bul. soc. chim. România* 12, 90-4 (1930).—In an attempt to settle the question as to whether paraffin, resins and asphalt are present in petroleum in colloidal or in true soln., petroleum of different types from 5 localities were subjected to ultrafiltration. Vulcanized rubber membranes of 0.04 mm. thickness and a vulcanization coeff. of 1.7 were used in the Brukner app. under a pressure of 150 atm. Two membranes were used that gave an effective filtration surface of 25 sq. cm. The resins of petroleum which are separable by 70% EtOH were found to be ultrafilterable. All the paraffin in the petroleum used was ultrafilterable but the asphalt, both hard and soft, was removed quantitatively by ultrafiltration. This behavior indicates that the asphalt is colloidal but that the resins and paraffins are not. The asphalt retained by the ultrafilter was dissolved in C_6H_6 and cyclohexane and was again retained by the filter. Several passages of an oil through the same filter removed the resins and left the pure dry asphalt, which dissolved readily in that portion of the petroleum that was ultrafiltered. This expt. shows that the asphalt exists in petroleum as a lyophilic colloid. ODEN E. SHEPPARD

Contact filtration of bright stocks. V. MATUSEVICH *Azerbaidzhanskoe Neftyanoe Khoriyisko* 1930, No. 12, 64-75.—A mixt. of long residuum from Dossor (85) and Makat (15) crude oil was treated with acid and in 2 stages with clay. Substitution of a NaOH wash for the 1st stage clay treatment resulted in the formation of emulsions and in high ash content of the finished oil. Lab. and plant results showed that long residuum requires 7-10% and cylinder stock 8-12% of 96-98% H_2SO_4 , which is applied as a fine shower, with air agitation. Long residuum is best treated at 30-35° and cylinder stock at 45-55°, since at higher temps. the sludge discolors the oil. The time of contact is detd. by testing the oil on a glass plate and discontinuing agitation when large lumps of sludge begin to form (usually 20-45 min.). The main body of sludge is sepd. from the oil at once and the remainder is allowed to settle for 15-48 hrs. in settling tanks at not over 50-55° (heating even with steam coils is detrimental). Coagulation of the sludge with water, NaOH soln., clay, water glass, etc., fails to give good results. Operations are controlled by color tests (2% naphtha soln. of oil in a Duboscq colorimeter). Clay should be applied to acid oil but the oil should be free from sludge. With long residuum a mixt. of clay (10% by wt. of oil), pretreated with acid and water (1:3), is mixed with the oil and circulated through a Foster-Wheeler pipe still (construction and operating details are specified) and a mixing tank where the temp. should reach 165° (max. permissible temp. in pipe still is 170°). When the temp. reaches 90-100°, 5-10% more clay is added (total 15-25%). If all of the clay is added at the beginning, foaming is excessive. This clay treatment requires 6-8 hrs., but lab. results show that 40 min. contact gives the best color. The use of superheated steam shortens the time of treating

but the product is impaired. Loss of oil to clay is about 1-1 by wt. Similar results are obtained with cylinder stocks but more clay is required. Clay is sepd from the oil on a filter press of the frame type, this is unsatisfactory because of loss of time. The following results were obtained in treating long residuum

98% H ₂ SO ₄ , % by wt.	—	6	9	12	15
Color of acid oil, mm	—	41	55	70	80
Sp. gr.	0.9182	0.9055	0.9046	0.9022	0.8997
Flash point	Practically unchanged				
Conradson C, %	3.70	1.68	1.48	1.16	1.07
Color after clay treatment, mm	0.5	50	68	87	91

V. KALICHEVSKY

Dielectric constants of petroleum and its products. V. ZAUZE *Azerbaidzhanskoe Neftyanoe Khozjalstvo* 1930, No. 12, 88-98.—The dielec. const. of petroleum and its products increase with the increase in sp. gr., b. p. and mol. wt. of oils. The same relation holds for paraffins. Contrary to the conclusions of Pécheux (*C. A.* 21, 109). The dielec. const. is very closely equal to $(n_D)^2$. For paraffins, the const. increases with the increase in m. p. The temp. coeff. of the const. is neg. for petroleum products. Its abs. value it increases with the d except for the paraffins, for which it shows a decrease. The temp. coeff. of paraffins is 2-3 times smaller than that of liquid mineral oils. The dielec. const. of naphthenic acids increases with increase in sp. gr., b. p., surface tension, index of refraction and mol. wt.

V. KALICHEVSKY

Absolute viscosities of petroleum products from Grozny fields. L. SELSKII and E. SICHIGLOVA *Azerbaidzhanskoe Neftyanoe Khozjalstvo* 1930, No. 12, 130-1

V. KALICHEVSKY

Cold settling of surakhani crude oil. A. VELIKOVSKI and S. NIKONTOVA *Azerbaidzhanskoe Neftyanoe Khozjalstvo* 1930, No. 10, 100-11.—Satisfactory lab. results were obtained.

V. KALICHEVSKY

Ligroin and gasoline content of straight-run and cracked distillates from Grozny crude oils. L. SELSKII *Azerbaidzhanskoe Neftyanoe Khozjalstvo* 1930, No. 12, 124-9

V. KALICHEVSKY

Many products from Burman crude oils. G. ECLOFF. *Oil and Gas J.* 29, No. 30, 40, 123 (1930).—The situation and production of the fields are discussed and the products obtained are described.

J. R. SIKOVA

Analysis of petroleum and its distillates for reducible substances and adsorbable matter by means of the polarographic method with the dropping mercury cathode. B. GOSMAN and J. HEVROISKY *Trans. Electrochem. Soc.* 59 (preprint) 23 pp. (1931), of *C. A.* 24, 3419.—The electrocatalytic method with the dropping Hg cathode and automatic polarographic recording of the current voltage curves has been applied to the estn. of (1) electrolytically reducible substances and (2) adsorbable (surface-active) matter in petroleum and its distillates. The analysis of electroreducible substances is carried out by shaking the petroleum with a soln. of NH₄Cl and submitting the aq. ext. to polarographic examn. after the atm. O₂ has been expelled by H₂ or by the addn. of a sulfite. About 0.005% of reducible substances may be detected in this way. Many sorts of petroleum were found to contain acidic substances, probably naphthenic acids, which may be extd. by aq. alk. solns. and which possess a prominent anticatalytic action in the reaction between O₂ and sulfite. These anticatalysts suppress maxima occurring on current voltage curves at pos. potentials, thus indicating their neg. charge and strong adsorbability. The content of adsorbable substances is detd. by adding the petroleum or its soln. in MeOH from a microburet to an acidified MeOH soln. of CuSO₄. The max. on the current voltage curve, due to the latter soln., is hereby depressed and the diln. at which a half suppression is reached is taken as characteristic for the adsorbability in the petroleum ext. Naphtha from different sources has been found to contain the least adsorbable substances (diln. no. 10-400), then gasoline (400-2000) and lubricating oil (3000-13,000). Crude oil is most adsorbable with diln. nos. up to 41,000. C. G. F.

Knocking and effect of antiknock compounds. J. LORNTZEN *Z. angew. Chem.* 44, 130-6 (1931).—The causes of knocking were investigated in combustion expts. made in a cylindrical chamber with mixts. of air with pentane and hexane. Variations in the proportions of the mixts. showed that the reactions are different when knocking occurs and when it does not. Change of position of the spark in the chamber proved that knocking sets in when an incomplete combustion is propagated through the chamber and a pressure wave is reflected from the wall, causing an explosion of the partly oxidized products. The flame is propagated about twice as fast under knocking conditions as

otherwise Turbulence increases the combustion velocity In the exptl arrangement used, knocking occurred in a turbulent gas mixt only at pressures above 8 atm The addn to the fuel of antiknock compds, lead tetraethyl and iron carbonyl, had no effect When the combustion products of the same antiknock agents were introduced into the combustion chamber previous to the ignition of the fuel, it was impossible to produce knocking below 5 atm pressure, and when turbulence was produced at the same time, no knocking was observed, although the initial pressure was raised to 9 atm

D F BROWN

Benzene and the evolution of ideas on motor fuels. BRUNSWIG *J usines gas* 54, 569-76(1930)—Current theories on the causes and prevention of motor knock, the methods of refining fuel benzene, gum formation and the action of inhibitors, and methods of testing for gum forming tendency are discussed K H ENGEL

Knock rating of motor fuels. Note on experiments carried out in 1930. ANOV *J Inst Petroleum Tech* 17, 69-71(1931)—From the results of a 2nd series of knock-rating tests participated in by the research labs of the Anglo American Oil Co, the Anglo Persian Oil Co and the Asiatic Petroleum Co, it was agreed that the bouncing pin was the best means of measuring knock intensity if the engine was small enough for it to function properly The Sub Comm of the I P T agreed upon the octane no for expressing knock ratings in the future The results from the older types of testing engines can be compared with those from the Homing engine only by working out the correlation for each type of fuel separately

EMMA E CRANDAL

Gases from crude oil cracked in the vapor phase. M B MARKOVICH and V V. PRIGULEVSKI *Neftyanoe Khozaystvo* 18, 425-44(1930), cf *C A* 25, 405—Fuel oil was cracked in the vapor phase and the gas produced was sep'd into 3 fractions (a) permanent gas, (b) Blau gas, i e, liquid at 15° under 100 atm pressure and (c) light gasoline These fractions contained, resp, in % by wt, H 24, 0, 2, 0, 0, satd hydrocarbons 49 0, 29 6, 3 0, C_2H_4 25 2, 17 1, 0 0, C_2H_6 12, 21 3, 4 0, butylene 2 0, 0 0, 35 0, divinyl 2 0, 6 6, 18 0, vapors of higher unsatd hydrocarbons 0 0, 2 5, 0 0, N_2 , CO_2 , O_2 , etc, 3 0, 0 4, 0 0 The work was undertaken to det the possibility of using cracked gas in the manuf of synthetic chem compds to replace natural fats, for prep glycols, perfumes, etc.

A A BOEHLINGK

The vapor-pressure curves of motor spirits. P MEYER *J Inst Petroleum Tech* 17, 42-68(1931)—The Ramsay-Young boiling-point law, deduced from the Clausius-Clapeyron equation, may be expressed in the form $\log P = K_1 - K_2(t/T)$ Vapor pressures of a large no of air free motor spirits were detd, in Andrew's app, and by the methods of Rhodes and McConnell, Beistle and Prather, Tizard and Marshall and Bridgeman, Aldrich and White All the results fell approx upon the graph represented by B, A and W's equation, in which the consts of the Ramsay-Young equation are evaluated $\log P = 6.76 - 3.83(t/T)$ Two methods were developed for calcg the true hydrocarbon vapor pressure from the apparent vapor pressure to avoid the exptl difficulty of removing dissolved air from a mixt of volatile hydrocarbons By the first method, the apparent vapor-pressure curve is regarded as made up of the true vapor-pressure curve plus an air pressure curve The 2nd method utilizes the apparent molal heat of vaporization for calcg

EMMA E CRANDAL

Vapor-phase treatment of cracked gasolines in the Vickers cracking unit. B P FRADEIN and D A STROM *Neftyanoe Khozaystvo* 18, 445-8(1930)—Gasoline obtained in the Vickers cracking unit is evapd again and passed in the vapor phase through fuller's earth filters The yield of the final gasoline amounts to 91 1%, while 0 2% of fuller's earth is consumed, based on 11 regenerations During this process the fuller's earth is continuously heated by flue gas, the temp being kept at 230-5°. The untreated gasoline has a Stammer color of 3 2, gum 0 5%, no corrosive action, unsatisfactory doctor test, and a sp gr of 0 734 The treated gasoline has a color of 2 1, 0 3% of gum, satisfactory corrosion and doctor tests, and a sp gr of 0 737

A A BOEHLINGK

Destructive hydrogenation in bomb-polymerization in cracking. S A KISS *Ind Eng Chem* 23, 315-8(1931)—The initial retardation of the decompn in destructive hydrogenation, carried out in a closed bomb, is explained by the Poynting effect which abstracts the "activated molecules" from the liquid phase and transforms them into "non activated molecules" Formulas are developed for the kerosene yield and the gas and gas plus-gasoline yield, also for the amt. of polymerization

D F BROWN

Cracking unit of Jenkins. B CHIRKIN *Azerbaidzhanskoe Neftyanoe Khozaystvo* 1930, No 10, 67-92—A complete description is given of the units bought from the Graver Co, including detailed drawings, dimensions, guarantees, etc.

V K

Design of liquid phase cracking equipment. P RYABUKH *Azerbaidzhanskoe Neftyanoe Khozaystvo* 1930, No 12, 76-82—If the cracking rate doubles for every 10°,

the cracking rate (C_m) for the temp interval $T_1 - T_2$ is $C_m = 14.5C_1/(T_1 - T_2)$ ($1.074^{T_1 - T_2} - 1$), where C_1 = cracking rate at T_1 . The length of the heating coil (L) is the sum of the length required to heat the oil (L_1) and the length necessary to supply the heat for cracking (L_2), called as follows: $L_1 = L_1' q_1' / q_1 n \theta = L_1 q_1$, where q_1 = cal used in heating 1 g of oil to the cracking temp., q_1 = cal absorbed in cracking 1 g of oil, n = the velocity of the oil in pipes in m/sec., θ = heating time in sec. At temps. below 450° L_2 is not over 2% and may be neglected but at 450° it may reach 25% of the total length. The temp drop in the reaction chamber under adiabatic conditions is $T_1 - T_2 = q_1(1 - \alpha) + q_2/(1 + \beta)S$, where q_2 = heat lost to the surroundings, α = the fraction cracked before the oil reaches the chamber, β = the fraction recycled; S = the sp. heat of the oil between T_1 and T_2 . For isothermic conditions $t_1 = T + (Z/GS)(Z - 1)$, $t_2 = T + Q/GS(Z - 1)$, where t_1 = the temp of the gases entering the heating coil, t_2 = the temp of the gases leaving the heating coil; Q = the heat absorbed in the reaction chamber, G = the wt. of gases used per hr; S = the sp. heat of gases, $Z = e^{Q/GS\Delta t_m}$, where e = the base of the natural log and Δt_m = the temp difference satisfying the relationship $\Delta t_m = Q/KH$, where K is the coeff of heat transfer between hot gases and the oil, and H = the surface of heating coils. Sample calcs are given.

V. KALICHEVSKY

Heat balance in cracking reactions. V. KJOSKILOV AND P. MARKOSOV. *Azerbaidzhan'skoe Neftovoe Khimicheskoe* 1930, No. 12, 83-7.—The work of Ryabukh (preceding abstr.) is criticized and a method recommended for calcy. cracking reactions is outlined.

V. KALICHEVSKY

Present methods for investigating gasolines. N. KALANTAR. *Azerbaidzhan'skoe Neftovoe Khimicheskoe* 1930, No. 12, 55-63.—Various analytical methods are described.

V. KALICHEVSKY

Refining of cracked gasolines with clays from Surakhani. I. GIDENBURG AND A. MARKARYANTZ. *Azerbaidzhan'skoe Neftovoe Khimicheskoe* 1930, No. 10, 93-R.—Clay alone gives unsatisfactory results but might be of value in combination with acid or hyposulphite.

V. KALICHEVSKY

Perfect flexible method for treating [gasolines and naphthas]. C. O. WILSON. *Oil and Gas J.* 29, No. 31, 24(1930).—A method has been perfected by the Union Oil Co. of Calif. for economically treating all grades of gasoline and naphtha by a continuous acid treatment followed by contact with clay. This company also uses the heat from the recycle stock of their Cross stills to rerun the treated pressure distillate.

J. R. STRONG

Determining gum contents of gasoline. OSCAR C. BRIDGEMAN AND ELIZABETH W. ALDRICH. *Oil and Gas J.* 29, No. 37, 42, 199, 200, 201(1931).—The Cu-dish method gave on the av. much higher results than either the use of a glass, silica or porcelain dish in an atmosphere of steam or the air jet method with glass dish and air blowing on the surface of the liquid at the rate of 300 ml per sec. On high gum gasolines the glass-dish method gave higher figures than the air jet, but for low gum gasolines they are about the same. The values decreased rapidly from zero air flow to 300 ml per sec and from there on did not change. Inert gases (CO_2 and N_2) gave somewhat higher results than those obtained with an atm of air. The methods thus far proposed have not been correlated with gum deposit in an engine. It is felt that a lab-flow method can be devised that will simulate engine-manifold conditions and that can be used as a preliminary criterion for the choice of a method to be employed in the detn. of gum content. This all seems to hinge on the true gum content, hence a method must be found to det. the true gum content.

J. R. STRONG

Determination of the gum content of cracked spirits. W. LITTLEJOHN, W. H. THOMAS, AND H. B. THOMPSON. *J. Inst. Petroleum Tech.* 16, No. 84, 684-95(1930).—This method of detg. gum, decided upon for tentative use in the laboratories of the Anglo-Persian Oil Co., is the outcome of a series of expts. in which the effects of varying many factors were tested. A 25-cc sample of gasoline is evaporated in a hemispherical glass dish 7 cm. in diam., on a steam bath, with the aid of a jet of air at 0.5 in. pressure. If the residue in the dish is not dry at the end of 1 hr. it is allowed to remain on the bath 15 min. after it is visually dry. Drying is continued in an oven at 150° for 1.5 hrs. If the gum exceeds 11 mg in wt., a new detn. is made upon a sample just small enough to bring the quantity of gum under the 11 mg limit. Results for vapor-phase cracked gasolines are more affected by the change from the old method of drying at 100° for 1 hr. than are those of gasolines formed by liquid phase cracking.

EMMA F. CRANDAL

Action of inorganic refining reagents on alkyl sulfides in naphtha. P. BORGSTROM AND J. C. McINTIRE. *Ind. Eng. Chem.* 23, 321-3(1931).—Expts. were run on the removal of alkyl sulfides dissolved in various samples of naphtha. The sulfides used were

ethyl, propyl, isopropyl, alkyl, butyl, isobutyl and sec-butyl sulfide. The reagents used were silica gel, mercuric acetate, mercuric chloride, solid mercurous nitrate, Zn sulfide, cupric sulfide, ferrous oxide, ferric oxide, cupric oxide, cobaltic oxide, Ni oxide, yellow mercuric oxide, Pb oxide, Al oxide, Zn oxide, chromic oxide, basic Cu carbonate, ferric nitrate, stannous chloride, Zn chloride and cupric chloride. The results show that the lower sulfides are more readily removed by mercuric acetate than the higher ones. The removal of the isobutyl and isoamyl sulfides was about the same as that of the normal compounds. More sec-propyl and sec-butyl sulfides were removed than the corresponding normal compounds. The action of mercuric chloride varied greatly. An appreciable difference was noted with mercurous nitrate according to the naphtha used. This reagent appears, however, to be the best available for the removal of alkyl sulfides.

D. F. BROWN

Influence of water and steam in the rectification process. K. KOSTRIN *Azerbaidzhanskoe Neftyanoe Khozylstvo* 1930, No 10, 99-103, cf C A 21, 1178. —Condensation of steam in the rectifying columns and not the steam itself has a harmful effect on rectification.

V. KALICHEVSKY

Origin and refining of sulfur oils. W. FRIEDMANN *Oil and Gas J.* 29, No 31, 32, 106, 107(1930) cf C A 24, 5989

J. R. STRONO

Practical aspects of modern dehydration. M. E. GARRISON *Petroleum World and Oil Age* 27, No 12, 67, 71, 72, 75, 77, 79(1930)

LEMA E. CRANDAL

Determining the congealing point of fuel oil. N. I. CHERNOZHUKOV AND A. M. GUTZAIT *Izvestiya Teplolekhn. Inst. (Trans. Thermo-Tech. Inst. Russia)* 1930, No 3, 33-5. —Put the fuel oil into a 200-cc flask (about half full) provided with a thermometer, stir the oil at a rate of 80-100 revolutions per min. on a water bath which is 3-5° cooler than the congealing temp. of the fuel oil. When the oil reaches the bath temp. transfer it to a test tube provided with a thermometer which has been brought to the bath temp. Insert the test tube in a second tube, and place the assembly in a cooling bath the temp. of which is kept 6° below that of the fuel oil. The congealing point is found by tilting the test tube at an angle of 45° at each degree.

A. A. ROBITLING

Sapropelites from the river Barsas (Siberia). N. M. KARAYAYEV AND J. B. RAPAPORT. *Izvestiya Teplolekhn. Inst. (Trans. Thermo-Tech. Inst. Russia)* 1930, No 3, 59-60. —Three types of deposits were found near the river Barsas: (1) dark-yellow, very dense, in layers, very difficult to grind; (2) dark gray, in layers, the layers becoming more distinct after drying, they can easily be sepd. and disintegrate into scales on grinding; (3) a piece of wood rot with bark on it. These sapropelites contained moisture 1.2-3.5%, ash 10-20.8%, S 0.6-1.0% and volatile matter 60.5-80.6%. The elementary analysis of the combustible part gave: C 76-79.6%, H 7.3-9.6% and N 0.23-0.72%. The calorific value was 8150-9100 kg. cal. A low-temp. carbonization in a Fischer retort yielded 38.27-57.21% semicoke, 33.81-50.00% tar; water and the gas and loss amounted to 7.83-10.83%. The coke so obtained was dense and was good only as a low grade com. fuel. The tar contained 2.35-3.81% phenols, 2.35-3.00% acids, 0.51-1.2% bases and 87.9-93.5% neutral components of 0.8137-0.8972 sp. gr. The tar distn. produced 16.7-25.4% b. below 200°, 35.9-43.6% b. below 275° and 56.1-64.1% residue b. above 275°. The gasoline fraction had a sp. gr. of 0.7604-0.7750 and an I no. of 136.7-142.3. It had an unstable color which, after treatment with H₂SO₄ of 1.84 sp. gr., washing with caustic soda and water and redistn., was stabilized, the loss through treatment amounting to 21.65%. The original gasoline fraction contained 21.50-31.5% of uncat. compds. and 9.3-15.9% of aromatic compds. Its sp. gr. after treatment and redistn. was 0.7679. The kerosene fraction was also unstable in color. It contained 13-14.11% phenols, traces of bases, 85.89-87% of neutral components and had a sp. gr. of 0.817-0.8030. The gas oil distillate was semi-solid (b. 280-350°) and contained up to 0.68% of paraffin, caked on the tar. In a destructive hydrogenation of 70 g. of the gas oil and 10 cc. of gasoline, with 2 g. of Al₂O₃ as catalyst, 67.5% of liquid product was obtained. The heating was continued for 4 hrs. at 415-25°. The original pressure was 70 atm.; the final and max. pressures were 65-75 atm. and 200 atm., resp. An insignificant amt. of C and gas was formed. The liquid had 33.15% b. below 170°, 16.31% b. 170-230°, 14.29% b. 230-300° and 5.28% b. 300-350°, the residue was 18.06% and the loss 5.89%. Expts. carried out with Al₂O₃ and Al₂O₃ + CuO by heating the mixt. for 2-3 hrs. produced similar results. A destructive hydrogenation of the residue b. above 280° in the presence of Al₂O₃ and CuO yielded 33% b. below 175°, 13.25% b. 175-260°, 11.45% b. 260-300°, 14.37% b. 300-340° and 27.93% of residue (including loss). All fractions were unstable in color. Cracking of the 280-350° fraction at 405-50° for 1.5 to 2.25 hrs. at a pressure of 18-62 atm. yielded 51.85-77.7% of a product b. up to 280° and 22.3-32.45% of residue. The I value varied between

62.38 and 35.94. The gasoline and kerosene fractions, although off color, are suitable for com. purposes because they contain some aromatic compds and unsatd compds.

A. A. BOEHLING

Oil-field waters of Alberta and Saskatchewan. W. P. CAMPBELL. *Trans. Can. Inst. Mining Met.* 32, 316-32 (1929).—The water problem in oil borings is discussed. A high chloride content generally indicates deep water; absence of chlorides and presence of sulfates and H carbonates indicate surface water.

B. C. A.

The making of low pour point oils. E. R. LEDERER AND E. W. ZUEBLIN. *Oil & Gas J.* 29, No. 32, 64, 101 (1930). cf *C. A.* 25, 1367.—There are 5 important methods for making an oil with low pour point: (1) solvent processes; (2) cold pressing; (3) cold settling; (4) reentrifuging; and (5) Weir process. Solvent processes have attained little importance as yet because of the possible loss of solvent. These processes make an oil equal to the best com. products. Cold pressing is applicable to overhead oils only, and depends entirely upon crysto. of the wax into needles rather than plates. By careful control of the distn. and pressing in 2 stages it is possible to make an oil whose pour point is 0°. Cold settling consists in diluting the stock with about 70% of light naphtha, chilling and settling. The wax must be amorphous, not cryst. The viscosity of the oil must be sufficiently low to permit the wax to settle out. The oil must have a low sp. gr. for the highest possible gravity differential between soln. and wax. Before centrifuging, the oil is slowly chilled in soln. with 70% of light naphtha, at a rate of 5° per hr, to about -60°F. Truly cryst. wax cannot be sepd. by centrifuging. The Weir process consists in filtering cold with the aid of Hyflo-Super Cel. One lb. per gal. is added to the diluted oil, and the cooled mixt. is forced through a filter. The Hyflo-Super Cel. is abrasive and cannot be handled with pumps.

J. R. STRONG

Gasoline and lubricants by hydrogenation. R. T. HASLAM AND W. C. BAUER. *Oil and Gas J.* 29, No. 37, 34, 165, 166 (1931).—Hydrogenated oils are, in general, equal or superior to all others in the relationships that are the usual indications of quality. In the White motor, oil consumption was least for the hydrogenated oil and with the exception of one other oil the C deposit was lower. Tests in the Mack motor were more severe, but results were still favorable to the hydrogenated oil. The hydrogenated gasolines are very sensitive to addn. of tetraethyl lead, have low S and gum, and are equal to C₆H₆ for blending purposes. Hydrogenated gasolines are unlike most cracked gasolines in that the heavier fractions have the best antiknock qualities.

J. R. STRONG

Turbine oils. AVON. 5th All Russian Thermo-Tech. Congress. *Izvestiya Teplo-tekhn. Inst. (Trans. Thermo-Tech. Inst. Russia)* 1930, No. 1, 66-9.—Emba turbine oil distilled from the Konstantinovskii refinery after being used for 6 months in a turbine caused the bearing temp. to rise up to 70°, a great amt. of sludge composed mostly of carbon and asphaltene being pptd. The same oil when overrefined became too viscous and the amt. of acid formed which was able to cause corrosion was quite appreciable. The same over refined oil with water being admitted during the test run in a steam turbine produced, after 4 months' use, a great excess of sludge. Oil of the above type but slightly acidic and contg. an insignificant amt. of sludge changed very little in the presence of water, although it did not stand up in the de-emulsification test, the latter fact being explained by the formation of Cu and Fe soaps, while the same oil under normal conditions worked quite satisfactorily. A sample of Baku turbine oil was satisfactory under normal working conditions. To regenerate used turbine oils they may be centrifuged to sep. out water and sludge, while oils high in acid require refining. The lubricating system of the steam turbine should be blown out with steam to remove impurities. Considerable data concerning turbine oils before and after use are given.

A. A. BOEHLING

Make asphalt by vacuum reduction. D. H. BANCROFT. *Oil and Gas J.* 29, No. 36, 30 (1931).—Asphalts of improved quality are being made in the El Dorado, Ark., plant of the Lion Oil Refining Co. by reducing the bottoms in a high vacuum. Heavy Smackover crude oil is processed through a combination dehydrator, topping unit and vacuum reducing tower. In addn., this unit makes a lubricating distillate, gas oil and gasoline. Asphalt can also be made from low-gravity Smackover fuel oil. The lower the abs. pressure the lower will be the ductility and the penetration.

J. R. STRONG

Improving crude-oil asphalts. A. I. VORONOV AND N. I. LOGVINOV. *Neftyanoe Khoz. i Stroitel'stvo* 18, 449-60 (1930).—Road asphalts meeting the various standard specifications were prepd. from paraffin base, mixed base and asphalt base fuel oils by distg. the oils under high vacuum or with steam or refinery gas, followed by blowing with air. The oil sepd. in distn. was suitable for prepg. lubricating oils and cylinder stocks or it could be used for cracking. Careful distn. was required to avoid the conversion of tars

into asphaltenes, which increase the brittleness of asphalt. An extensive bibliography on asphalt is appended. A. A. BOEHLINGK

Solubility of asphalts and pitches in various, less common, organic solvents. EMIL J. FISCHER. *Asphalt u. Teer* 30, 1332-4 (1930).—See C. A. 24, 2589. K. H. ENGEL

The chemical nature of bitumen from Petchora. N. D. ZELINSKII. *Compt. rend. acad. sci. U. R. S. S.* 1930A, No. 14, 351-2.—The bitumen is a black, viscous mass mixed with vegetable residues and possesses a rather sharp, characteristic odor. A sample was extd. with benzene. After the benzene was distd. off, the residue was distd. under vacuum (11 mm.). The distillate was a yellowish oily liquid which distd. under ordinary pressure without decomn. between 190° and 305°, and had the following compn.: C = 86.48% and H = 13.34%, $d_4^{20} = 0.8390$. The rather high d indicates the predominance of heavy hydrocarbons of cyclic nature. The residue from the first distn. was distd. under high vacuum and the remaining mass was a hard tar of the following compn.: C = 85.00% and H = 10.64%. This tar constitutes the main part of the Petchora bitumen. L. JACOVLEFF

Carbonization (of wood) in modern portable furnaces. ANTONIO BERNAD. *Instituto Forestal de Investigaciones y Experiencias* 2, No. 5, 68-80 (1929).—The disadvantages of the old method of wood carbonization (slowly burning open piles of wood) are such as to cause the gradual disappearance of this method. The use of portable metal furnaces improves the quality and yield of charcoal, eliminates the necessity of const. attention to the burning, prolongs the working season, prevents interruption by adverse weather conditions and supplies a source of power for running machinery. The details of construction and operating expenses are given. The Spanish woods used for the manuf. of charcoal are *Pinus sylvestris*, *P. halepensis* and *Eucalyptus globulus*. H. E. S

Solid CO_2 [from oil well gas] (MARTIN) 18. Residue of sulfonic acids from oil-refining (Fr. pat. 694,234) 18. Temperature-indicating devices for oil-cracking apparatus (U. S. pat. 1,792,086) 1. Destructive hydrogenation of pastes of coal and oil (Brit. pat. 336,610) 21. Liquid products from coal, etc. (Fr. pat. 694,192) 21. Apparatus for extracting gases from oils (Fr. pat. 693,985) 1. Apparatus for supplying fire-extinguishing foam to oil tanks (Brit. pat. 337,221) 18. Dialyzing membrane [for purification of mineral oils] (Fr. pat. 693,007) 18.

KOETSCHAV, R.: *Erdöl und verwandte Stoffe*. Dresden and Leipzig: T. Steinloppf. M. 8.

MIRCHELL, H. V.: *Fuel Oils and Their Applications*. An Introductory Treatise on the Sources, Classification and Production of Fuel Oils, with Notes on Collodial Fuels. 2nd ed., revised by A. Grounds. London: Sir Isaac Pitman & Sons. 176 pp. 5s.

WAGNER, H.: *Taschenbuth der Farbenkunde*. Stuttgart: Wissenschaftliche Verlagsgesellschaft m. b. H. 240 pp. M. 6.50.

Purifying petroleum oils. KARL T. STECK (to Standard Oil Development Co.) U. S. 1,791,941, Feb. 10. Oil is treated with a liquid mixt. of SO_2 and SO_3 .

System for treating petroleum oils with reagents such as acid and adsorbent earths. SAMUEL J. DICKY and ERNEST W. ROTH (to General Petroleum Corp. of Calif.). U. S. 1,792,003, Feb. 10. The oil under treatment is repeatedly and successively mixed with a mixt. of finally purified liquid and reagent sludge, the sludge is drawn from the mixts. formed, and successive treatments are effected with mixts. contg. successively higher percentages of unexhausted reagent. An arrangement of app. is described.

Desulfurizing petroleum oils. FRANCIS M. ROGERS (to The Standard Oil Co.). Can. 309,099, Mar. 3, 1931. Straight run, low-boiling petroleum distillates are desulfurized by passing the vapors through adsorbent clay maintained at a temp. around 630°F., whereby a substantial part of the combined S is disengaged as H_2S .

Distillation apparatus with entrainment separators for vacuum distillation of petroleum oils. ALEXANDER C. SPENCER (to Standard Oil Development Co.). U. S. 1,791,940, Feb. 10. The vapor stream from a vacuum still is cooled without substantial condensation. The direction of flow of the stream is then suddenly changed, its linear velocity decreased and the vapors are passed to a condensing zone. Various details of app., including an entrainment separator, are described.

Hydrocarbons. MELANTI A.-G. Fr. 694,682, April 29, 1930. Carbonaceous or hydrocarbon materials are converted to liquids of lower b. p. by prep. a mixt. of the substance to be treated with a colloidal soln. of a metal such as colloidal Zn, and treating

the mixt. under pressure and heat, e. g., by injecting it in the atomized state with the aid of H₂ or gases contg. H into appropriate reaction vessels.

Purifying hydrocarbons. GENERAL TECHNICAL CO., LTD. Fr. 695,004, Aug. 2, 1929. Fractions of various hydrocarbons (petroleum, tars, etc.), particularly fractions of low b. p. (benzene, gasoline) are distd. by adding to the fraction a non-miscible liquid or a mixt. of non-miscible liquids, the b. p. of which is lower than that of the lightest sulfured compd. in the mixt. The mixt. is afterward distd. at the boiling temp. of the non-miscible liquid introduced so that the sulfured impurities are sepd. from the hydrocarbons and a distillate is obtained free from S. Water, alc., certain higher alcs. and ketones or halogen derivs. of hydrocarbons, alcs., aldehydes or ketones may be used.

Decolorizing and stabilizing hydrocarbon mixtures. I. G. FARBENIND. A.-G. (Walter Rousinsky and Karl Smeykal, inventors). Ger. 516,270, Aug. 17, 1928. The mixts. are treated with a small proportion of an aldehyde- or ketone-sulfoxylate at atm. or raised temp. An alc. may be added during the treatment, and afterward distd. off. Thus, crude lubricating oil may be treated with an equal amt. of MeOH and 1% of Na formaldehyde sulfoxylate and the alc. then removed.

Removal of naphthenic acids from hydrocarbon oils. FRANCIS M. ROGERS (to The Standard Oil Co.) Can. 309,100, Mar. 3, 1931. Naphthenic acids are removed from hydrocarbon oil by vaporizing the oil and passing the vapors upward through a tower maintaining molten NaOH in the lower part of the tower and producing a small amt. of reflux in the upper part of the tower to prevent entrainment of Na compds.

Hydrocarbon oil conversion. GUSTAV ECLORF and JACQUES C. MOORELL (to Universal Oil Products Co.) U. S. 1,791,618, Feb. 10. After heating oil to a cracking temp. in a heating zone such as a coil of pipe in a furnace, vapors thus formed are subjected to dephlegmation and the dephlegmated vapors are subjected to catalytic treatment such as with Ni on pumice and the treated vapors are condensed and the distillate and uncondensed gases are collected, the uncondensed gases are scrubbed, passed through a desulfurizer and then admitted to a flowing stream of dephlegmated vapors prior to the subjection of the latter to the catalytic treatment. App. is described.

Hydrocarbon oil conversion. WM. R. HOWARD (to Universal Oil Products Co.). U. S. 1,791,622, Feb. 10. A body of oil is maintained at a cracking temp. under super-atm. pressure in a zone of vapor evolution, vapors are taken off and subjected to a primary reflux condensation and the reflux condensate is returned to the zone of vapor evolution and there heated to the prevailing temp. but out of direct contact with the other material in such a zone (suitably by a reflux still within the main still) and vapors are evolved from the reflux condensate and these vapors are combined with the uncondensed vapors from the reflux condensation, and the combined vapors are subjected to secondary dephlegmation and condensation. App. is described.

Converting heavy hydrocarbon oils into lighter products. LESTER KIRSCHBAUM (to Universal Oil Products Co.) U. S. 1,791,566, Feb. 10. A stream of oil is heated to a cracking temp. (suitably in a pipe coil in a furnace) and then transferred to a reaction zone, and the temp. imparted to the oil stream is controlled in accord with variations in the thermal condition of the oil stream as it is transferred to the enlarged reaction zone, to maintain it at a substantially uniform predetd. temp. App. is described.

Purifying light hydrocarbons produced by destructive hydrogenation of oil, coal, etc. R. SCOTT and IMPERIAL CHEMICAL INDUSTRIES, LTD. Brit. 335,885, May 31, 1929. Purification is effected by passing the material, while still in the vapor state and under a pressure of at least 10 atm., through a solid purifying agent such as bauxite or fuller's earth. App. and various details of procedure are described. Regeneration of the purifying material may be effected by treatment with air at 400-500°.

Purifying light oils recovered from coal gas, etc. J. A. SILAW (to Koppers Co.) Brit. 336,636, Nov. 6, 1928. Materials such as light oils recovered from coal gas and distillates from tar and petroleum are purified by washing with acid and then with an alk. soln. of Na phenolate in 2 stages, first in such quantity that most or just all of the acid present is neutralized and practically all the phenols liberated from the phenate dissolve in the oil, and second to effect neutralization and remove the phenols and any remaining acid. H₂SO₄ is usually employed. There are thus obtained from the first alkali washing solns. contg. practically no phenols, and from the second solns. of Na phenolates practically free from sulfates, sulfonates, etc., and these latter, with or without sepn. and recovery of the phenols, may be used for purifying gas or for softening water, etc.

Light hydrocarbon oil distillation and purification. D. A. HOWES, R. SCOTT and IMPERIAL CHEMICAL INDUSTRIES, LTD. Brit. 336,615, July 22, 1929. Kerosene and light hydrocarbons produced by destructive hydrogenation, cracking or straight distn.

are distd at a suitable temp below 300° in the presence of gaseous NH_3 and in the absence of inorg acids, to improve color and restrict tendency to gum formation Various examples are given.

Apparatus with superposed drums and heating fines for distilling hydrocarbon oils. JACKSON R. SCHONBERG (to Standard Oil Development Co.) U S 1,791,937, Feb 10

Cracking hydrocarbons. PAUL VEROLA Fr 694,495, July 31, 1929. In cracking hydrocarbons by heat treatment, the gases produced in the process are reintroduced partially or wholly into the liquids before, during or after the cracking thereof has taken place, whereby the properties of the residual oil are improved

Cracking hydrocarbons. SOC D'ÉTUDES ET RÉALISATION "ERÉAL" (SOC. ANON.) Fr 694,744 Apr 29, 1930 Gaseous products rich in hydrocarbons are decomposed to obtain valuable by products by passing them through an incandescent mass The speed of the current or the height of the layer traversed or both are regulated so that at the exit of the mass temps are obtained which do not exceed that necessary for a product of good quality The temp of the mass is about 1000–1200° and of the exit gas 600–700°.

Cracking hydrocarbon oils. FRITZ HOFMANN and CARL WULFF U S 1,791,562, Feb 10 See Brit. 301,395 (C A 23, 4059)

Cracking oils. I. G. FARBERIND A-G Brit. 337,046, July 26, 1929 Metal compds of 1,3 diketones sol in hydrocarbon oils, such as the W, Mo, Cr, V, Mn, U, Nb, Zr, Ce, Sn, Ti, Al, Cu, Zn, Co or Ni compds of acetyl-, propionyl-, and butyryl acetates, are used, singly or in various mixts with each other or with other suitable substances, as catalysts in converting into lower h p compds materials such as mineral oils, tar oils or oils obtained by the destructive hydrogenation of carbonaceous materials Various details and examples are given, and the process may be carried out at 300–600° with or without high pressure or use of H

Cracking oils. A. P. SACHS (to Petroleum Conversion Corp.) Brit. 336,360, Oct 6, 1928 In a cracking process in which the heat for the conversion is supplied by a heated carrier circulated cyclically, the H_2S which accumulates in the gas is washed from it by spraying the gas with water after it has been cooled to condense the vapors produced (and preferably after it has passed through the circulating blower) Various details of app and procedure are described

Cracking mineral oils. S. SEELIG Brit. 336,822, Nov 7, 1928 See Fr. 684,456 (C. A. 24, 5477).

Apparatus for cracking oil. GUSTAV EGLOFF and HARRY P. BENNER (to Universal Oil Products Co.) U. S. 1,791,617, Feb 10 Oil is passed in a single passage through a pipe coil in which it is heated to a cracking temp, and a valve-controlled connection is provided for passing heated oil in liquid phase from the coil into either of the opposite extremities of an enlarged horizontal vapor chamber sepd. from the coil, and means, such as pipes of restricted diam, is provided for increasing the velocity of the oil as it reaches points near the entrances to the vapor chamber The vapor chamber communicates with a condenser, and the app is arranged for conducting the oil treatment under superatm pressure Cf C A 24, 4385

Apparatus for distilling and cracking hydrocarbons. NIKOLAUS MAYER. Austrian 120,428, July 15, 1930 Addn to 118,255 (C A 24, 4927)

Oils. I. G. FARBERIND A-G (Carl Wulff, inventor) Ger 513,814, April 23, 1927. Aryl carbonates, formed by treating the phenols with COCl_2 , are used as transformer and switch oils

Refining mineral oils. RICHFIELD OIL CO OF CALIFORNIA Fr 695,077, May 5, 1930 See U S 1,790,622 (C A 25, 1374) Fr 695,078 describes a process for refining hydrocarbons by vaporizing them and bringing them into contact with a soln of at least one metallic salt, e g, a salt of Zn

Refining oils and waxes. COSIPAGNIE PROVENÇALE DE RAFFINAGE DE CORPS GRAS MINÉRAUX Fr. 694,469, July 27, 1929 Mineral oils and waxes are refined by dissolving them in naphthalene decahydride and treating with H_2SO_4 in the usual manner The decahydride is afterward recovered by distn

Atomizing heavy oils. KARL MODL Austrian 120,725, Aug 15, 1930 The atomization of heavy oils in a gas stream is facilitated by introducing a readily vaporized fuel into the gas stream The method may be applied in connection with internal combustion engines

Fractional condensation of oil vapors. HARALD NIELSEN and BRYAN LAING. Ger 513,613 Oct 30, 1928 The vapors are passed through a rotary washer under which various fractions condense as the temp. falls App is described

Condensing and separating apparatus suitable for use in steam distillation of

mineral oils. W. M. W. HOLLAND (to Standard Oil Co. of Ind.) U. S. 1,792,164, Feb. 10. Structural features of a baffle and gravity sepn. app.

Decolorizing oils such as vapor-cracked gasoline or other mineral oil or distillate. N. A. MIJN BOUTEN, C. LUTTERMAAR-SCHAPPEL BOUTEN. Brit. 331,952, Nov. 3, 1928. Decolorization is effected by solns. of alkali metal hydrides in substantially water free org. solvents miscible with the oil such as MeOH, EtOH, BuOH, glycerol or acetone, and small quantities of reducing agents such as Zn, Al, aldehydes or polyhydric phenols may be present. The alkali may be recovered by dissolving the still residue in water and calcining. Cf. C. 4 25, 1374.

Refining gasoline. AUGUST P. BYEREGAARD. U. S. 1,791,521, Feb. 10. After removing water from gasoline the dry gasoline is mixed with a high test alc. and this mixt. is treated with dry caustic alkali.

Refining cracked benzene. OTTO ZECHETMAYR. Ger. 513,870, June 20, 1928. The benzene is led through layers of porous material contg. hydrides of alkali or alk. earth metals. The layout is described.

Use of mercury cyanide as an "antidetoxant" with motor fuels. A. J. DUNCAN. Brit. 335,928, Dec. 4, 1928. Several mixts. are described.

Separating paraffin from oil. ROBERT E. MANLEY (to Texas Co.) U. S. 1,792,231, Feb. 10. A wax bearing lubricating-oil stock is mixed with an added more volatile carrier oil such as a light asphalt base lubricating oil which is less volatile than kerosene, and the carrier oil is then distd. off together with the paraffin wax under less than atm. pressure. An arrangement of app. is described.

White petrolatum. HUGH F. GALLAGHER (to Standard Oil Development Co.) U. S. 1,791,924, Feb. 10. A petrolatum yielding material is treated with fuming H_2SO_4 in the presence of a heavy, normally liquid hydrocarbon material such as lubricating oil distillate to obtain a liquid product contg. suspended sludge, a readily settling sludge such as asphaltic bottoms is added to this product so that the suspended sludge is pptd., and petrolatum is obtained from the residual reaction mixt. App. is described. Cf. C. 4 24, 493.

Fractionating tar, oils, etc. J. KARPAT. Brit. 331,945, June 19, 1929. See Fr. 691,360 (C. 4 25, 1003).

Tank and distributing nozzle system, etc., for treating bottom settlings or cut oils with water, steam and soda ash. W. M. YANTIS. U. S. 1,791,623, Feb. 10. Structural features.

Distilling bituminous materials. WERCHEN-WEISSENFELSER, BRATKOWKEN A. G. Ger. 513,506, Oct. 6, 1928. Details are given of a vertical retort with an endless band or chain for circulating the bituminous material.

Treating bituminous sands, etc. KARL A. CLARK (to Governors of the University of Alberta) U. S. 1,791,797, Feb. 10. In sepg. bitumen from bituminous sand, silt or clay the material is mixed with a reagent such as Na selenate giving an alk. reaction on hydrolysis, and the mixt. is introduced into a body of hot water to effect sepn. of the bitumen from the sand, silt or clay, a water sol. inorg. compd. of a multivalent metal such as $CaCl_2$ is introduced into the hot water to coagulate the reagent and the silt or clay and the bitumen is recovered from the surface of the water. App. is described.

Continuous production of dispersions of thermoplastic material such as bitumens, gums and waxes. GEORGE A. BROWN (to Bennett, Inc.). U. S. 1,792,007, Feb. 10. Various details of app. and operations are described.

Lubricating and insulating oils. I. G. FARBENING, A. G. (Mathias Pier and Ernst Hochschwender, inventors) Ger. 516,316, June 1, 1927. See Fr. 688,102 (C. 4 24, 1499).

Lubricants. STANDARD OIL DEVELOPMENT CO. Fr. 694,225, April 19, 1937. Lubricants are obtained from a heavy hydrocarbon oil contg. impurities of a gummy or resinous nature by eliminating all the non-cryst. impurities from the oil, e. g., by distn. and subsequent treatment with H_2SO_4 , and treating the purified oil with a gas rich in H under pressure and at a temp. at which decompn. is relatively slow, e. g., 371-426°.

Reconditioning lubricants such as crank-case oil from internal-combustion engines. M. CZARNY and F. C. FROLANDER. Brit. 336,302, June 6, 1929. The lighter fluid impurities are absorbed by a wick and are evapd. from the wick in an app. which is described and which may be arranged to operate continuously in connection with an engine.

Apparatus for cleaning used oil by heating and filtering. FIRMA HERMANN BAYMANN. Ger. 516,183, Feb. 22, 1927.

Composition for removing carbon deposits and rust from internal-combustion engine cylinders or other surfaces. GEORGE R. WAGNER (one-half to Wm. L. Smith)

U S 1,792,052 Feb 10 Kerosene 1 gal, lubricating oil 1 qt, "lubricating graphite" 10-15 oz, ether 3 oz and spirits of camphor 3 oz

Turpentine oil, pine oil and rosin from coniferous woods. FRANCESCO C PALAZZO

U S 1,792,392, Feb 10 See Fr 676,859 (C A 24, 3f10)

Turpentine. JEAN B H FONGA Fr 36,835, April 18, 1929 Addn to 668,493 (C A 24, 1737) Turpentine from maritime pine is treated with a solvent, the distn temp of which is below that of the turpentine, and the resinic acid normally present is pptd by adding lime or other alk earth base The solvent is recovered and a turpentine of reduced acidity and absolute purity is obtained

23—CELLULOSE AND PAPER

CARLETON E CURRAN

Construction of the crystalline part of cellulose. H. H MARK AND K H MEYER *Cellulosechemie* 11, 91 100(1930) cf C A 23, 5077 24, 1500—Cellulose interference to x rays can be explained by a monoclinic structure, it leads to an atomic arrangement which is illustrated, and to micelles with a length of about 500 and width of about 50 A U The micelle surface reactions and the "permutoid" reaction are taken into account. According to the Hludon rule, the principal valence chain model makes possible the estn of the rotatory power of various cellulose derivs which is in agreement with practical knowledge A method for x ray comparison of cellulose preps is given and minor structural differences are discussed C A BRAUTLECHT

Cellulose in the light of the x-rays. WM DRAGO *Cellulose* 1, 80-3, 110-3 (1930)—See C A 24, 2287, 5151 C A BRAUTLECHT

Structure of ramie cellulose as deduced from x-rays. O L SPONSLER AND W H DORR *Cellulosechemie* 11, 186-97(1930)—On the basis of new exptl evidence with x rays, S, and D deduce a structural formula which appears to agree with known chem and phys characteristics Three dimension models, employed for comparison with location of atoms in the space lattice, showed the following characteristics (1) The structure-forming units are arranged parallel and lengthwise in continuous chains. (2) The individual chains are arranged in rectangular order of dimensions 6.10×5.40 A U (3) Individual units in each chain repeat every 10.25 A U (4) One can assume individual cells to have the dimensions $6.10 \times 5.40 \times 16.25$ A U. (5) In the individual cell in every chain there are 2 C₆ units (6) In the elementary cell the atoms are so arranged that a no. of planes appear whose space intervals have a simple numerical relationship to 10.25 , namely, there are planes present which lie in intervals of 5.15 , 3.40 , 2.58 , etc It is concluded that the proposed mol structure agrees with the x ray analysis and apparently with all phys and chem characteristics; the cellulose structure is built up of glucose units, the amylose oxide ring formula agrees best with the x ray findings, the glucose units are built up into endless chains through glucosidic condensation, combination between units is of a 1 1 and 4 4 type bridge; the 1:4 bridges which are characteristic of cellobiose do not exist in the structure of ramie cellulose, the chains are rectangular 6.10×5.40 A U separately arranged with the greatest length along the diagonal direction, a group of 8 glucose units is the simplest unit which the cellulose structure can represent, agreeing with the crystallographic unit with axes $a = 10.80$, $b = 12.20$, $c = 10.25$ A U, the ramie fiber is a hollow cylinder in whose walls the crystallographic units are so arranged that the diagonal of the 6.10×5.40 space always takes a tangential position, the cellulose structure is stabilized, in length by primary valence forces which hold together the glucose units, transversely by the secondary valence forces of O atoms, the suggested structure permits the elasticity of the fiber in the direction of its length, explains in length and breadth, various coeffs of elasticity and allows for computation of swelling phenomena, the structure indicates, in conformity with Irvine's methylation results, that the OH groups attached to the 2nd, 3rd and 6th C atoms are open for substitution It is shown that ester formation is possible, without change of the fiber structure, only by spreading apart the longitudinal chains, the fiber structure can thus be weakened, and, if the introduced groups are large, the secondary valence forces can be sufficiently weakened so as to lose the ability to retain the fiber-like arrangement, since ramie cellulose is like other cellulose fibers it is probable that similar celluloses have a similar structure, the presence of xylose units in several celluloses can be carried through to the accepted belief that oxidation of glucose units leads to gluconic acid, which on subsequent decarboxylation yields xylose; and, that these changes would not appreciably change the suggested arrangement of the original fiber structure. C. A. BRAUTLECHT

Acetolysis of bamboo cellulose. S OGURI AND M NARA *Chem News* 142, 134-5 (1931) — See C A 25, 1072 G. G.

Bamboo II Bamboo lignin. K. SISIDO *Cellulose Ind* (Tokyo) 6, 167-70 (1930) Abstracts 29-32 (in German); cf. C. A. 25, 1669 — Bamboo is especially suitable for a study of lignin formation in vegetable material, because the shoots grow extraordinarily quickly. Shoots 5.8 m long were divided into 3 parts, each part being then dried and pulverized, and the dry powder extd exhaustively with alc. CaH_2 and dried *in vacuo*. Lignin was removed with 72% H_2SO_4 and its methoxyl no detd by the Zeisel method. The methoxyl nos for the samples were also calcd on the assumption that this no represents lignin. Both calcd and observed values show a decrease from the base to the tip of the shoots, but good agreement is not found. It is considered probable that in the samples more remote from the base of the shoot the whole of the methoxyl is not fixed to lignin, and it is concluded that the lignin is methylated gradually according to the growth of the bamboo, or combines with the residue which contains Me in order to give the higher values detd. Some properties of bamboo lignin have been compared with those of wood lignin, and it is found that they behave similarly. Thus C H, ash and methoxyl contents are similar in amt. to those from spruce wood lignin. Also both spruce wood and bamboo chips give a green coloration on immersion in concd HCl , while, on sepn., the lignin gives a grayish violet coloration before, and a salmon pink after washing with boiling water. This suggests that bamboo lignin also has a heterocyclic structure. The distribution of O in the lignin as hydroxyl, methoxyl, etc., is discussed. B. C. A.

Integral relationship of cellulose to difficultly soluble xylan in the structural substance of red beech (*Fagus sylvatica*). I. E. SCHMIDT, K. MEINER, K. NEVROS AND W. JANNESBUR *Cellulosechemie* 11, 49-68 (1930) — Previous observations by Schmidt, *et al.*, have shown that (a) dil aq ClO_2 affects lignin markedly and can be used to identify lignin or its fragments, (b) lignin does not undergo exchange with the carbohydrate part of the cell wall (c) oxidation products of lignin which are insol in water can be removed through salt formation with alkaline salts such as Na_2CO_3 , or action with H_2SO_4 in form of NaHSO_4 , phenols or pyridine, (d) the unsatd part of the cell wall (lignin) can be removed quantitatively, (e) the polysaccharides of the cell membranes are converted into water sol. forms, especially d galactose and carbohydrates which behave as derivs. of d galactose such as L arabinose, which differ from the carbohydrates of cell membranes that are built up from d glucose and its conversion products such as d mannose. Investigation of the skeleton substance of the wooded cell walls of cryptogams and phanerogams showed that COOH-contg polysaccharides always participated in the building up of skeleton substance, a synthesis best explained by the ester theory. The above evidence led S, *et al.*, to investigate the quant. relationships of single polymer units of the skeleton substance and the manner of their combination. In red beech it was shown that the skeleton substance treated with 0.04 to 0.2% NaOH quant. loses an easily sol. xylan, while ester like combinations of acetyl groups are split off. The residue from the treatment with 0.04 to 0.2% NaOH is entirely acid free and consists of cellulose and xylan. The skeleton substance contains two xyans, an easily sol. (I) and a difficultly or less sol. one (II). II can be sepd from the acid free mixt. of cellulose and I by soln. in 5% NaOH at ordinary temp. On the assumption that cleavage products of cellulose, such as result by 5% NaOH action, are dissolved simultaneously with xylan, there is a stoichiometrical relationship of one part xylose anhydride ($\text{C}_5\text{H}_8\text{O}_4$) in II to 3 glucose anhydride ($\text{C}_6\text{H}_{10}\text{O}_5$) in the cellulose. This observed regularity of skeleton substance in red beech is dependent upon the kind of after treatment during the breaking up of the wood with ClO_2 , when an excess of OH ions is prevented (which can be the cause of specific side reactions), also on age, standing, time of cutting, etc. It was found that the proportion of cellulose to II was 3:1 for red beech independent of age or geographical origin. Thus for the first time a marked mass-weight association is shown between two colloids. This leads to a view of the compn. of wood which apparently contradicts the incorporation theory based on a mixt. of cell membrane cellulose, pentosans, hexotans and lignin. Part I is devoted to the method of hydrolysis with ClO_2 and Na_2SO_3 with reference to improvements of the method. With pH limits of 6.8-7.5, after ClO_2 treatment, lignin can be removed quant. by Na_2SO_3 . Part II is devoted to the action of NaOH of different concns. (0.04 to 17%) on the skeleton substance of red beech. With concd alkali solns the work was carried out under N. The carboxyl-contg part of the skeleton substance is removed completely by 0.2% (0.05 N) NaOH. Polymeric carboxylic acid and polymeric xylose anhydride is extd. simultaneously with the hydrolysis of acetyl groups. The easily sol. I is sol. in 0.2% NaOH; the difficultly sol. II is attacked completely only by

5% NaOH. When 3% or 1% NaCl soln. is added to the alkali the plotted curve for % wt loss against concn of alkali takes a different course, being the same, however, for the 3 as for the 1% NaCl. Salts, other than NaCl, also affected the wt loss. The const. wt. loss with 5-7% NaOH contg. NaCl, implies that concns of 3% NaOH and upward cause soln. of II and part of the cellulose. Five % NaOH in presence of 3% NaCl seps xylan quant. from cellulose. Thus from a mixt. of cellulose and (II), 5% NaOH contg. NaCl dissolves 21.39% xylan and leaves 78.61% cellulose, and upon dividing the % quantities by the equiv. wts., 132 for xylan and 162 for cellulose, the quotients 1.3 are obtained.

C. A. BRAUTLECHT

Integral relationship of cellulose to difficultly soluble xylan in the structural substance of red beech [*Fagus silvatica*]. II. E. SCHMIDT, W. JANDEBEUR AND K. MERNEL. *Cellulosechemie* 11, 73-9 (1930), cf. preceding abstract.—Red beech wood meal treated with 0.25% ClO_2 in water oxidizes lignin, yielding products which can be removed by treatment with salt solns. having a p_H of 6.8. Na_2SO_3 serves to dissolve lignin (1) by yielding Na ions to form salt like sol. lignin derivs. and (2) by H_2SO_3 from Na_2SO_3 to form sol. products. At a p_H value of 6.8 (a higher p_H affects the polysaccharides) salts other than Na_2SO_3 , such as Na_2CO_3 , Na_3PO_4 , or K xanthate can be employed under the same conditions of temp. (50-60°) leading to the same stoichiometrical relationship between cellulose and less sol. of the two xyans present. Repeated alternate treatments of SO_2 , in form of 4% NaHSO_3 soln., and water-sol. phenols such as resorcinol, with ClO_2 completely removes lignin from the cell constituents. Polysaccharides are not attacked by this treatment and even after 400 hrs. treatment with ClO_2 , the ratio for delignified material of 3 mols. cellulose to 1 mol. less soluble xylan remains the same.

C. A. BRAUTLECHT

The relation of the viscosity of cellulose ester solutions and temperature. III. E. BERL, H. UMSTÄTTER AND E. KARRER. *Z. physik. Chem., Abt. A*, 152, 284-94 (1931), cf. C. A. 24, 5151.—The previous equation is found to hold for nitrates or acetates of cellulose. The viscosity does not change after heating and cooling if no acid is present. The Berl-Büttler equation holds with sufficient accuracy for such viscose materials. An improved app. is described.

V. F. HARRINGTON

Some remarks on acetylcellulose. FRITZ OHL. *Farbe u. Lack* 1931, 51, 64.—A review. To test for the presence of cellulose sulfoacetate O dyes the sample with a 1% aq. soln. of methylene blue. Particles contg. sulfoacetate are dyed a darker blue than acetate particles.

G. G. SWARD

Changes in the viscosity of cellulose acetate-acetone solutions. C. G. RITTENHOUSE AND C. J. STAUD. *Paint, Oil & Chem. Rev.* 91, No. 9, 9-10 (1931).—The viscosity of cellulose acetate-acetone solns. increased faster than the concn., an indication of a colloidal system.

G. G. SWARD

Chemistry of pectin and its relation to the formation of incrustations of cellulose. I. F. EHRLICH. *Cellulosechemie* 11, 140-51 (1930); cf. C. A. 24, 65.—A summary of the chemistry of pectin, its general and com. phases and its discovery by Braconnot in 1825 to the present being considered. Much attention is devoted to pectic acids isolated from the com. pectins sepd. from beet sugar waste and cult. citrus fruits. Attention is also given to galacturonic acids, their isolation, relationships and characteristics. II. *Ibid.* 161-70.—Enzymic transformations of pectin and possible formation of lignin from pectin are discussed. Cyclic tetragalacturonic acids *a* and *c* can be opened by pectin ferments, especially by *Penicillium glaucum* and *Perisporiaceae*, the latter are easily obtained from sugar-beet pulp. Taka-diastase acts on a partly neutralized soln. of galacturonic acids *a* or *c* at a p_H of 4-6 in the course of a day to yield tetragalacturonic acid *b*. The new *Perisporiaceae* enzyme named "pectolase" by E. is obtained in especially active form from exts. of old cultures of isolated *Perisporiaceae* molds, in dry form. With 5% pectolase, galacturonic acids *a* and *c* could be converted into galacturonic acid *b* in a few min. and this could be totally converted into cryst. *d*-galacturonic acid in 2 or 3 days. The behavior of enzymes on tetragalacturonic acids leads E. to regard them as original components of pectin. The enzymes from molds which have the most intense action on the middle lamellae also act most intensely in the hydrolysis of tetragalacturonic acids. E. suggests that the tetragalacturonic acids which have been previously obtained only by acid hydrolysis of pectin are also intermediate products of the original pectin mol. which have been degraded by enzyme action. The mechanism of the action of pectolase on pectin involves, under specific conditions, (1) the splitting off of MeOH, (2) the reaction of the liberated pectic acids with Ca salts to form an insol. Ca salt of galacturonic acid *a*, which seps. as a gelatinous coagulum. In studying metabolic changes in highly lignified plant parts, E. examd. flax in the same manner as sugar beets. In old flax pectin, because of a hydration, the tetragalacturonic

acid ring is opened so that the galacturonic acid moles in the pectic acid are held together as members of an open chain. Very marked differences are shown by flax pectin as compared with other pectins, with regard to the alc components of its hydratopectins. Whereas hydratopectins from beets, oranges, etc., yield 20-25% of araban contg only arabinose, the fraction of flax hydratopectin sol in 55-70% ethanol is a complicated hexopentosan mixt. of arabinose, xylose, galactose and fructose. On evapn. of the alc soln of this crude hexopentosan, 25% of a resinous substance resembling lignin is obtained. After washing with cold water, and repeated soln in NaOH and pptn with HCl, a dark-brown powder results which is insol in water and acids but sol in alkalis and NH_4OH and shows no trace of carbohydrate. Its reactions and analysis show that furan nuclei are an integral part of it. E. concludes that the original substances of lignin in plants are really pectins and that the substances isolated from partly lignified flax straw represent intermediate products in the change from pectin to lignin. This resin and the hexopentosan from sugar free flax straw were firmly held to the parent substance, because prolonged extn with 70% ethanol did not remove them. In comparing the lignin (resinous product above) with other lignin preps E. states that (a) wood contains practically no pectin, (b) the larger part of the lignin is found in the middle lamella of wood, analogous to the pectin of the new nourishing tissues of young plants, (c) it appears that the methoxyl and acetyl content is derived from pectin, (d) all kinds of pectins, in addn to galactose, are characterized by a high content of pentoses or pentose like groups such as arabinose, xylose and tetragalacturonic acids (which can be regarded as carbonyl pentosans) and (e) it is very probable that a conversion of pectin into lignin is caused by chem. and enzymic processes during growth and aging. The change can be represented by the equation: $\text{C}_{11}\text{H}_{19}\text{O}_{11} \rightarrow \text{C}_{11}\text{H}_8\text{O}_5 + \text{CO}_2 + 10\text{H}_2\text{O} + 6\text{O}$.

C. A. BRAUTLECHT

Viscose. XXXIV. Effect of chlorination on the properties of cellulose. S. MASUDA AND J. MURAKAMI. *Cellulose Ind* (Tokyo) 6, 173-4 (1930); *Abstracts* 33-4. — Previous work by the authors has shown that the properties of sulfite pulp are improved by chlorination of the material at 15°, and similar results are now found for cotton. The Cu no. of the purified cotton cellulose used was reduced from 5 to 4 as a result of chlorination for 15 min at 21° and washing with water, under standard conditions and to 3 by similar chlorination for 30 min.

B. C. A.

Recent developments in the preparation of viscose and viscose silk. HANNS SCHMIDT. *Chem Fabrik* 1931, 73-7, 83-90, 97-100. — A review of improved methods and machinery with suggestions of possible simplifications. German and foreign patents are cited and 52 literature references are given.

J. H. MOORE

The enzymes of barley malt. VI. Fermentative decomposition of viscose silks. H. FRINGSHEIM AND E. TIRLO. *Cellulosechemie* 11, 100-2 (1930), cf. Otto, C. A. 23, 4955. — On fermentation certain small specific hydrolytic differences were observed; also that there does not appear the slight hydrolysis of viscoses pptd. in the presence of strong salt concn. in comparison to viscoses coagulated with H_2SO_4 alone. Viscose silks pptd. with $(\text{NH}_4)_2\text{SO}_4$ and then fixed with dil H_2SO_4 are least decompd. and differences in behavior with other enzymes are probably due to differences in absorptive properties of the surface layers of the viscose filaments.

C. A. BRAUTLECHT

Silk fibroin as raw material for artificial-silk preparation. HANNS SCHMIDT. *Z angew Chem* 44, 83-4 (1931). — A review.

H. H. MOSHER

Factors during spinning which influence the physical properties of rayon. II. PHILIP C. SCHERER, JR., AND ROBERT E. HUSKEY. *Ind Eng Chem* 23, 297-300 (1931), cf. C. A. 24, 3359. — For specific conditions in which only temp was varied, elongation and tensile strength show definite min and max.

FOSTER DEE SNELL

New bark [and sap] of wood of spruce, pine and red beech. C. G. SCHWABER AND K. E. NEUMANN. *Cellulosechemie* 11, 113-28 (1930). — Extns of sugar like substances, hydrolyses, detns of pentosans, extn of bark with $\text{Ca}(\text{OH})_2$ and $\text{Ba}(\text{OH})_2$, acid hydrolysis after swelling with NaOH soln and detns of resin (at wax content were made. Wood-forming saps were investigated. Bark can be easily removed in the spring because the cambium cells are very tender. In fall and winter, these cell walls are heavily incrustated and the bond between bark and wood is a strong one. Dil acids dissolve the incrustated substances in spruce and pine. Similar treatment has no action on the beech. Hot water extn showed only negligible quantities of sugar in the bark of all 3 trees, indicating that sugars played no important part in the formation of the incrustation. Results of hydrolyses with dil acids were different between the coniferous and deciduous barks. The conifers yielded more sugar from the inner bark than from sap wood, indicating the presence of easily hydrolyzable hemicelluloses in the barks of these trees. All representatives of hexosans, commonly found on wood hydrolysis

(except glucose), could be shown to be present, but no pentosans. The pentosans were therefore not attacked by the dil. acid. The bark contains more pentosans than the sap wood but the ratio is less than that of the hexosans. Since the bark-incrusting material is mostly hexosans, in coniferous woods, dil. acids effect a partial soln. of incrusting material and therefore a weakening of the bond between bark and sap wood. The supposed mother substances of the hemicelluloses, carbohydrates and pectin substances were detd., also the ether sol. and substance volatile with steam, also ash and adsorbable tannins. In addn., the presence of important org. groups such as acids, alcs., aldehydes, phenols, etc., was tested for by pptn. and color tests. The total residues of all saps were small. The ash in residues was high because of the presence of appreciable quantities of org. salts. From the ash analysis of autumn wood and bark, especially from the Ca content, S. concludes that the mineral salts in the saps are stored in the bark, after the growing period. Substances adsorbed by hide powder were very high in pine and lowest in beech and yielded pos. tannin tests. Reducing sugar content of the saps was appreciable, being highest in the beech. These included hexoses, which were preformed in the sap. The content of total fermentable carbohydrates after hydrolysis was greatest in the beech. The pentose content of all the saps was remarkably small, but greatest in beech, although still small in ratio to hexoses. This suggests that the pentoses are derived from hexoses and that they are transformed in the same measure as produced. The ether exts. were chiefly resins. The fat and wax accumulation in the cambium layer of the beech in the autumn was not observed in the sap. Substances volatile with steam (etheral oils, terpenes, etc.) were found only in small quantity in the saps, chiefly in pine. Oxalic, succinic and citric acids could be sep'd and identified in beech as well as in pine sap. Lactic acid could be found only in beech. Tartaric, malic and acetic acids could not be found. Oxalic, succinic and citric acids could have been formed by simultaneous oxidation and splitting of carbohydrates in the growing process. Lactic acid is probably due to fermentation of the sap subsequent to extn. The acids exist free, in part, and yield an acid reaction in the sap. Neither LiOH nor MeOH could be detected. Coniferyl alc. was found in pine sap. Vanillin was found in the 3 saps, and it is assumed that coniferyl alc. is converted, in part, by oxidases or enzymes, into vanillin. Aliphatic aldehydes were absent. Guaiacol was present in very small quantity. The aromatic substances, which were found, all belong to the protocatechuic group and may be intermediate products in formation of lignin, although the quantity is less than that of lignin in the finished wood, especially in the conifers. This is addnl. support for the view that lignin is derived from the hexoses present in excess over the intermediate products (pectin, pentoses, pyrocatechol derivs.). Proteins are present only in traces and are probably derived from the protoplasm of the living cambium cells. C. A. BRAUTLEIGH

Apparent destruction of coniferous wood by larva of common beetle [*Hylotrupes bajulus*, L.]. R. FALCK *Cellulosechemie* 11, 89-91 (1930).—In studies to det. the changes in wood substance caused by insects inhabiting wood, F. exam'd fir wood attacked solely by this larva, dividing it into (1) the fine powder from the bore hole and (2) the adjacent wood. The air-dry wood contained 9.6% water, the borings 9.4%. The borings which had passed through the food tube of the larva were of the same color and in other phys. characteristics was similar to the pulverized wood, except as to degree of pulverization. Ash, cellulose, lignin, fat, wax, resin, pentosans and hemicellulose were det'd. in both the larva borings and sound wood powder. The lignin and resin contents were the same, total pentosan was practically the same, the larva borings contained 12.27% less cellulose. The hemicellulose from the larva borings contained 2.96% less glucose (equiv. to 2.67% less hexosan), 1.07% more pentosan in soln. and 0.76% less pentosan in the residue. H_2SO_4 (69%) at 10-12°, in a shaking machine for 2 days, removed 12.29% more cellulose and hemicellulose from sound wood powder than from larva borings. Conclusion: The larva of the common beetle removes about 21% cellulose and hemicellulose from the wood without changing the remainder. F. suggests that the lignocellulose complex, pre-existent in the original wood, is unaffected. C. A. BRAUTLEIGH

Apparent destruction of wood by larva of *Annobium* [common wood worm]. R. FALCK *Cellulosechemie* 11, 128-9 (1930), cf. preceding abstract.—Analysis of sound pine wood meal and of the finely disintegrated wood after digestion in *Annobium striatum* showed practically no difference between the 2 products as to fats and resins, alkali sol. material, pentosans and lignin. The insect apparently consumes about 9% cellulose, and, since the lignin content is the same, it is probable that the cellulose consumed is loosely, or not at all, combined with lignin in the original wood. Of the hemicelluloses a slight difference exists in the hexosan part, being 2.62% less in the worm

product (= 10% of the total hexosan fraction), the pentosans in the soln are 0.53% higher and in the residue 1.5% less. These results compared with those having to do with the common beetle are the same as to kind but differ in quantity. The analytical data lead to the conclusion that insects living in wood which do not yield a different colored residue utilize only a certain cellulose fraction, which is not combined with lignin or is only slightly combined.

C. A. BRAUTLEIGH

The visible digestion of wood for pulp. H. R. HAMMOND, *Chemist Analyst* 20, No. 2, 17-8 (1931).—A glassy liner digester capable of withstanding 150 lbs. pressure is shown.

W. T. H.

Some observations on digesting sulfate pulp. DAVID JOHANSSON *Svensk Pappers Tid* 33, 916-28 (1930).—Digesting in autoclave and testing the resulting pulp gave the following results: (1) The strength of the ground pulp increases with the increasing Roe Cl no., being max. at Cl no. 7. Ease of grinding decreases with increasing Cl no. (2) The max. removal of lignin occurs with a 25% Na₂S content. A lower Na₂S content raises the lignin content. The strength no. is lowest for soda pulp and highest for Na₂S pulp. (3) In spite of careful treatment and low final temp. during digesting it was impossible to make soda kraft pulp of the same strength as sulfate kraft pulp. Addn. of bleach liquor does not increase the strength. The lignin content is higher for soda pulp than for sulfate pulp for corresponding yields and appearance. (4) Shortening the chips down to 10 mm. does not affect the strength of the pulp. (5) Sulfate pulp from Douglas fir, goblolly pine and southern yellow pine from the U. S. compared with sulfate pulp from Swedish pine and spruce showed that the Swedish pulp, in spite of its shorter fiber, gave a stronger paper.

WILHELM SÖDERBLOM

Determination of the bleaching of sulfite pulp by potassium permanganate. H. J. SRENSK *Pappers Tid* 33, 928, 31 (1930).—In detg. the bleaching of sulfite pulp in the lab. by the Bjorkman no. method it was found that dry undelibered pulp gave varying results. It is necessary to delibered indirectly digested pulp before treating with KMnO₄. An improved method is described, involving dry pulp, delibered in H₂SO₄ soln., successive treatments with KMnO₄, 1 c (NH₄)₂(SO₄)₂, 6 H₂O, H₂SO₄, and KMnO₄, and calcn. in percent of CaOCl₂ by a special curve. Exact data as to quantities, times and temp. are given.

WILHELM SÖDERBLOM

Uses of paper in electrical apparatus. W. H. ANDERSON *Elec. J.* 28, 99-103 (1931).—See C. A. 25, 1083.

W. H. BOYNTON

Protecting wallpaper from fungi (Brit. pat. 330,241) 18.

Cellulose. HERBERT KÖRBER, Austrian 120,351, Jan. 15, 1930. Cellulose materials such as sulfite or soda cellulose are refined by treatment with hydrocarbons. The treatment may be effected in the hollander and the hydrocarbons may be used in the form of an aq. emulsion. Thus, unbleached cellulose in 15-20 times its wt. of water may be treated at 80° for 1 day with 10% of its wt. of hydrogenated naphthalene.

Cellulose. I. G. FARENHED A-G (Emil Hubert and Karl Wenbrod, inventors). Ger. 513,693, Oct. 11, 1927. Stable solns. of cellulose in CuO-NH₃ are obtained by adding catalyzers, especially CN compds., before or during the dissolving process to prevent deterioration of the soln. by the action of atm. O. The amt. of CN compd. added does not exceed 5%. Examples of CN compds. mentioned are K₄FeC₆N₄, KCN, HgC₂N₂, and K₂FeC₆N₄.

Bleached cellulose from materials such as wood. ERIC HAGGLUND U. S. 1,792,009, Feb. 10. Raw material such as beech wood is partly decompd. by boiling with alkali and the partly decomposed material is then treated in aq. suspension with Cl for a time sufficient to impart a uniformly brown color to the material, and after this, decompn. is completed by addn. of hypochlorite, by which intermediary lower Cl-contg. acids are formed which change the color from brown to yellowish white; the material is then washed and subsequently further bleached with Ca(OCl)₂.

Saccharifying cellulose. COMMERCIAL ALCOHOL CO., LTD., J. S. ARTHUR and R. GOGARTEN Brit. 337,017, April 20, 1929. Cellulose material which has a low water content or which is dried down to about 10% moisture content or less is treated with steam until it has a moisture content of about 20-40%, treated with HCl gas under cooling until decompn. is effected, and then, without removing the HCl, further treated with steam or a mixt. of steam and inert gas to saccharify the decompn. products (the vessel preferably being heated to 60-70° during the saccharification). Various auxiliary treatments are described.

Cellulose saccharification. COMMERCIAL ALCOHOL CO., LTD., J. S. ARTHUR and R. GOGARTEN Brit. 336,934, April 20, 1929. In the treatment of sawdust or other

cellulosic or closely related materials, the material is dried (suitably to a moisture content of about 10% or less) and is treated with steam and a volatile acid gas such as HCl or HF uniformly to distribute the acid and water throughout the material, and decomposition is then effected with the same acid under such conditions of low temp (suitably -10° to -15°) that no substantial formation of sugar occurs. Saccharification is then effected without removing the acid by direct steam heating or by indirect heating to 50 – 100° , followed by de-acidification and conversion of the sugars to monoses by treatment with steam under pressure or by boiling with acidified water. Brit. 336,935 relates to a generally similar procedure in which, however, the starting material may contain up to 50% of moisture.

Treating cellulose. CAMILLE DREYFUS. Fr. 694,879, May 2, 1930. The properties of cellulosic materials are improved by treating them with alk. liquids at a temp below atm., e.g. a 10% NaOH soln may be used at -10° . The product may be activated subsequently with AcOH or HCOOH with a view to esterification.

Molded products from cellulosic material. FREDERIC H. SIKS (to General Elec. Co.) U. S. 1,792,254, Feb. 10. A cellulosic material such as sawdust, bagasse or cornstalks having natural substances capable of yielding resinous compds is digested with NaOH , and H_2SO_4 is added to the digested mass to ppt. the dissolved resinous constituents; the material is filtered and the mass is washed until neutral, dried at about 80° and there is added to the dried mass about 5–40% of hydrated lime, the material is powdered and is molded under heat and pressure.

Cellulose derivative. SOC. LYONNAISE DE SOIE ARTIFICIELLE and PIERRE CIEVALET. Ger. 516,462, Dec. 23, 1926. A formylacetylcellulose is prepd by treating cellulose at a low temp with HCOOH in the presence of a small proportion of H_2SO_4 , then kneading the mass with AcOH contg more H_2SO_4 , and finally pptg the product with water or EtOH . An example is given.

Cellulose derivatives. LEON LILIENTHAL. Ger. 516,461, May 20, 1924. See Brit. 231,809 (C. A. 19, 3592).

Cellulose derivatives. HENRY DREYFUS. Fr. 694,881, May 2, 1930. Weighted or mordanted materials of, or contg, cellulose esters or ethers are prepd by incorporating a weighting or mordanting metallic radical, or an agent capable of pptg this radical, with the spinning soln, used for making the materials. The metallic radical may be used in the form of a thiocyanate. Examples are given of the use of compds. of Sn and Fe .

Nitrocellulose. IMPERIAL CHEMICAL INDUSTRIES, LTD. Fr. 694,623, April 28, 1930. Nitrocellulose is prepd by making, without disaggregation, a uniform cellulose having a high apparent d and giving limpid solns, by systematic subdivision of pulp in the form of plates or sheets to obtain pieces or tablets practically uniform, and nitrating these uniform tablets with a mixt. of H_2SO_4 and HNO_3 .

Alkali cellulose press. HANS HASELMANN. Ger. 513,863, May 8, 1927. Details of construction are given.

Cellulose esters. JOHANNES ALLES. Ger. 513,541, Mar. 20, 1928. Solns and plastic masses of cellulose esters are prepd by using esters of copal resin acids with mono- or dihydric alcs as solvents or softening agents. In an example, nitrocellulose, Congo copal benzyl ester and a solvent of the acetone, AcOAc or AcOAm type are mixed, and the product is diluted with spirits, C_6H_6 , etc. Other examples are given.

Cellulose esters. I. G. FARBENIND. A-G (Albert Gundlach and Theodor Becker, inventors). Ger. 516,250, Feb. 24, 1927. Esters of cellulose with higher homologs of AcOH are prepd by treating cellulose or its conversion products with chloroacetic anhydride and higher homologs of AcOH in the presence of a catalyst, e.g. H_2SO_4 or ClCH_2COOH . The soly. properties of the products can be improved by after-hydrolysis. Examples are given. Cf. C. A. 25, 1671.

Cellulose aliphatic acid esters. K. VERNER. Brit. 336,349, Aug. 15, 1929. In the production of cellulose aliphatic esters with preliminary removal of water from air-dried cellulose, before esterification, by treatment at temps up to 50° with sufficient aliphatic acid and anhydride (such as acetic, propionic or butyric acids and anhydrides) to complete the esterification reaction, the removal of water is facilitated by the presence of a small quantity of catalyst insufficient to promote esterification such as H_2SO_4 , 0.02–0.05% or SO_2Cl_2 , 0.05% (calcd. on the wt. of the cellulose) and a milling process is preferably used for the pretreatment. Various details and examples of procedure are described.

Apparatus for cellulose acetate production. JEAN ALTWEGG (to E. I. du Pont de Nemours & Co.). U. S. 1,792,059, Feb. 10. An app. is described comprising a column having an inlet adjacent to one end and an outlet adjacent to the opposite end, and pro-

vided in its interior with rotatable stirring devices of varying form, which increase in size from inlet to outlet according to the texture of the material to be treated' by these devices.

Apparatus for production of cellulose xanthate. M. DAWSONVILLE. Brit. 337,149, July 1, 1929. Numerous structural details are described of a rotatable app. through the journals of which CS₂ may be fed to the reaction chamber and heating or cooling agents similarly fed to a jacket around the reaction chamber. *CL. C A 24, 3646.*

Hydroxyalkylcellulose xanthates in manufacture of artificial threads, films, etc. L. LILJENFELD. Brit. 335,993, March 25, 1929. Hydroxyalkylcellulose xanthates such as those described in Brit. 335,994 (following abstract) may be worked up into threads alone or with other materials such as cellulose xanthate or alkali sol. derivs. of cellulose, gelatin, or softening or plasticizing agents, etc. (various examples of which are given). Numerous details of procedure are described for making products such as artificial threads and films, staple fiber, plastic masses, adhesives and cements, coatings and fabric dressings, artificial leather, transparent sheets, etc.

Xanthated hydroxyalkyl derivatives of cellulose. L. LILJENFELD. Brit. 335,994, March 25, 1929. Products of this type are produced by xanthating hydroxyalkyl derivs. of cellulose obtained either by the action of a hydroxyalkylating agent on cellulose in the presence of an alk. agent such as caustic alkali or by hydroxyalkylating alkali cellulose. Various halohydrins may be used as reagents and numerous examples, details and modifications of procedure are given. The final products may be isolated from the reaction masses or solns. (suitably after neutralizing with a weak acid such as HCl) by pptn. with MeOH or EtOH, Al salts, NH₄Cl or (NH₄)₂SO₄, NaCl, Na₂SO₄, NaHSO₄, CO₂, SO₂, H₂SO₄ or HCl or by dialysis. Catalysts such as salts of Cu, Ni, Ag, Zn or Fe may be added to the alkali cellulose or to the reactive must, and instead of NaOH or KOH there may be employed sulfonium hydrosulides such as trimethylsulfonium hydrosulfide. "Near conversion products" of cellulose such as cellulose hydrate, hydrocellulose or oxycellulose may be used in the process.

Transparent celluloid stable to light. SCHERING KAGLABATH A-G. Brit. 336,981, Oct. 5, 1928. Celluloid of otherwise usual compn. contains a small proportion of added ingredients capable of combining with O compds. of N to form colorless non injurious compds. such as urea or urethane.

Apparatus for treating viscose silk with liquids after spinning. ACME RAYON CORP. Fr. 694,971, May 3, 1930.

Textile spinneret. MINER M. ARSTIN (to Faustel Products Co.) U. S. 1,791,785, Feb. 10. A textile spinneret is provided with a hard resistant oxide film formed directly thereon to protect the spinneret against corrosion and wear.

Artificial fibers. I. G. FARBERVIND A-G (Harry Meyer, inventor). Ger. 513,664 Nov. 15, 1928. Mat viscose fibers are produced from viscose to which the water insol. liquid residues from the catalytic production of alc. have been added.

Device for connecting artificial threads to winding apparatus in the dry spinning process. J. E. PEDDER and COURTAULDS, LTD. Brit. 336,718, Aug. 21, 1929. Structural features.

Cellulose acetate yarn. CAMILLE DREYFUS and WILLIAM WHITEHEAD (to Camille Dreyfus). Can. 309,131, Mar. 3, 1931. Artificial textile products are produced by spinning a soln. of an acetone-sol cellulose acetate in a solvent mixt. contg. 93-97% of acetone b. 50° and 5-2% of MeOH b. 64.5°.

Cellulose acetate yarns. CAMILLE DREYFUS and WILLIAM WHITEHEAD (to Camille Dreyfus). Can. 309,130, Mar. 3, 1931. Artificial textile materials are produced by spinning solns. of acetone-sol cellulose acetate contg. less than 1% of moisture in a solvent mixt. of approx. 80% acetone and 35% deaerated alc. contg. substantially 93% of EtOH.

Artificial silk. J. P. BENBERG A-G. Ger. 517,018, Nov. 8, 1927. In the manuf. of artificial silk from ammoniacal copper solns. of cellulose by stretch spinning, the aq. pptg. bath contains a readily sol. gas, preferably a gas having a reducing action, e. g., SO₂ or H₂S. The bath may be preliminarily deaerated wholly or in part. *CL. C A 24, 3401.*

Artificial silk. I. G. FARBERVIND A-G (Hermann Schmidt and Emil Hubert, inventors). Ger. 516,370, June 12, 1928. Cuprammonium cellulose solns. are spun into pptg. baths contg. free CO₂ but no other dissolved gas.

Artificial silk. W. M. P. DREAFER. Fr. 693,923, April 15, 1930. See Brit. 328,627 (C A 24, 5490).

Artificial silk. RUSKE FURUSHO, YASUO USUMI and ASATUENON TODA. Fr. 694,285, Apr. 22, 1930. The thread of artificial silk after desulfurizing, etc., is brought

conditions that swelling of the cellulose deriv takes place Various details and examples are given.

Estification of wood. I G FAERBEND A-G Brit. 236,009, July 22, 1920. Acylation of wood such as beech, aspen, fir or pine is effected with an anhydride of an org acid such as acetic, butyric or benzoic anhydride, in the absence of inorg catalysts. The wood may be used in dry disintegrated form or in the form of sheets which may be superficially acylated, and may be initially dewatered and dewatered with solvents and treated with steam.

Treating bagasse fibers. ERNEST C H VALET (to Cellulosa Hemmer Valet, S A) U S 1,702,202, Feb 10. In order to render bagasse fibers suitable for paper manu, etc., they are treated with a clear said soln of lime and there is subsequently added a soln contg a sulfite such as Na_2SO_3 and the fibers are then treated in a soln of KOH, NaOH and a sulfite and are subsequently washed in a soap soln.

Sieve cylinder apparatus for dewatering cellulose pulp. J. STRINDLUND Brit. 330,870, June 11, 1929. Structural features.

Feltable bituminous pulps. CHARLES L KELLER (to Richardson Co) U. S. 1,702,094, Feb 10. Fibrous material is subjected to a "pre heating" to form a half stock and the latter is coned to bring it to an easily handled plastic stage, transferred to a mixer, premixed with added bituminous material and the "premixed" stock is transferred to a beater and subjected to heating to prepare the fibers for felting, and production of a product in sheet form.

Nitrocellulose from wood-pulp board. N. PICTON and IMPERIAL CHEMICAL INDUSTRIES, LTD Brit. 336,235, May 27, 1929. Wood pulp board is cut into substantially uniform pieces (suitably about one-eighth of an in. square) by a cutter which does not affect the agglomeration of the fibers and the pieces are treated with acid contg not less than 40% HNO_3 to produce a nitrocellulose of high d and uniform compn suitable for the manu of propellant explosives, blasting explosives, varnishes, celluloid, etc. Various details of procedure are described.

Tanks and associated agitating apparatus for treating paper pulp. JOSEPH E BOVN U S 1,701,705-6, Feb 10.

Dewatering paper pulp, etc. RICHARD EASTNER and HERBERT SCHMOLEA Austrian 120,000, Aug 15, 1930. Addn to 115,901 (C A 24, 2295). In the method of Austrian 115,901, the heating of the materials in the press is effected by passing an elec current through the materials, the metallic sieves acting as electrodes.

Dirt trap for paper pulp, etc. FRANK HASSMANN Ger 510,212, June 3, 1928.

Paper. I G FAERBEND A-G (Rudolf Richter, inventor), Ger 510,023, Aug 12, 1928. Ornamental or effect paper is prepd by addn of finely comminuted leather, colored or uncolored, to the paper stuff.

Paper stock. HARRY P BASSETT U S 1,702,351, Feb 10. Cotton linters or similar material is subjected to a combined slight heating and acid treatment with a weak acid soln such as a 0.05-1.0% soln of H_2SO_4 at a temp between 20° and 100° to effect partial hydrolysis of the material, and the latter is then treated with a weak alkali soln.

Recovering fibers from the waste waters of the paper, cellulose, etc., industries. GEORG WATZINGER Austrian 120,610, Jan 15, 1930. A gas under pressure is dissolved or emulsified in the waste waters, and the pressure is then released so as to liberate bubbles of gas within the liquid. These carry the suspended fibers to the top of the liquid, whence they are removed. App is described.

Paper-making machine (Fourdrinier type). BELFORT IRON WORKS Ger. 510,121, June 3, 1928. See Brit 268,180 (C A 22, 1238).

Suction box for paper-making apparatus, etc. GEORGE PEEBLES (to Selbec, Inc.) U S 1,701,535, Feb 10.

Doctor blade mounting for paper machines. L S JOHNSON and F. W. VICKERY Brit 338,554, May 14, 1929. Structural features.

Apparatus for reinforcing paper or other fabrics with unspun fibers. G E GINS (to American Reinforced Paper Co) Brit. 336,830, Nov. 8, 1928. Structural features. Printing and waxing paper webs in a single operation. HARRY C. COLE U. S. 1,702,414, Feb 10. Various details of app and operation are described.

Preparation for coating paper. SOC I F. LAUCKS, INC. Fr. 694,224, April 19, 1930. The prepn is made by incorporating a mineral substance with a protein in the liquid state derived from oleaginous seeds. The seeds are freed from oil and extd with a feebly alk salt such as Na_2SO_4 and the protein is pptd with lime.

Apparatus for coating paper with wax. A GASEN. Brit 335,968, July 2, 1929. Structural features.

Sizing for paper. ARTHUR W. BURWELL (to Alox Chemical Corp.). U. S. 1,791,860, Feb. 10. A material for use in sizing paper consists essentially of water-sol. alkali metal salts of normally liquid water-insol., petroleum insol., relatively high mol. wt. aliphatic and hydroxy-carboxylic acids such as those formed by oxidizing a paraffin-base fuel oil distillate and is capable of being retained in a paper in quantity up to 50% by wt. App. is described, suitable for oxidation of the oil.

Paper sizing. B. WIEGER (to Becker & Co. (1924), Ltd.). Brit. 335,902, July 1, 1929. Paper or pulp is sized with an emulsion prepd. by adding to the molten sizing material, while rapidly agitated, 1-3% of an aq. dispersion of a protective colloid (preferably alkali casein) and colophony is preferably employed as the sizing material, although there may be used other natural or artificial resins, solid or semi-solid hydrocarbons, waxes, fats, higher fatty acids or their mixts., with or without drying oils. Various details and examples are given.

Automatic device for regulating the drying of paper, etc., sheets. JOSEF MUCHKA. Austrian 120,248, July 15, 1930.

Drying cardboard, etc., by pressing. RICHARD PRÜCKNER. Austrian 120,700, Aug. 15, 1930. The press is housed in a casing having means for adjusting the temp. and pressure of the air in the casing. Operative features are described.

Water-resistant paper product. WM. C. LONCE (to Oswego Falls Corp.). U. S. 1,792,382, Feb. 10. There is added to the pulp sizing in the proportion of 3 parts rosin or rosin substitute and approx. 1 part wax or wax substitute, and a fixing or pptn. treatment is employed for the production of a final paper product contg. about 1% wax and not more than 3% rosin, which is suitable for bottle hood caps, etc.

24—EXPLOSIVES AND EXPLOSIONS

CHARLES E. MUNROE AND C. G. STORM

Report of Chief Inspector, Bureau for the Safe Transportation of Explosives and Other Dangerous Articles, for 1930. B. W. DURN. *B. E. Report* No. 24, 83 pp. (March 3, 1931), cf. C. A. 24, 2296—Only 15 minor accidents involving 2 injuries to persons and a total property loss of only \$71 resulted during the transportation of explosives in the U. S. during 1930. None of these accidents was with com. blasting or military explosives, but involved toy torpedoes, fireworks, etc. In comparison, inflammable liquids caused 602 accidents with 9 deaths, 4 injuries and loss of \$340,847, while acids and corrosive liquids caused 356 accidents, with 15 injuries and loss of \$21,057. Defective tank cars carrying H_2SO_4 and faulty packing of HNO_3 were the causes of most accidents with acids. The report gives statistical details and contains much information of interest to manufacturers, dealers and consumers of explosives.

C. G. STORM

Static charges in smokeless powder. H. E. NASH. *Army Ordnance* 11, 293-6 (1931)—To est. the danger of ignition from static electricity in a smokeless powder dryhouse, expts. were conducted in a special unit of one tray, 36" X 48" X 6", holding 100 lb. of powder, under varying conditions of moisture content of powder, and humidity, temp. and velocity of air current. Conclusion: (1) A moisture content of less than 1% is necessary for the existence of a measurable static charge. (2) A negligible amt. of static electricity exists on powder being dried in a heated air current. (3) Agitation of the powder causes larger static charges, but the max. charge is lower than that capable of causing ignition. With E. C. powder, the static found was less than $1/10$ that required to cause ignition.

C. G. STORM

Variation of $sp\,dt$ with loading density for different types of powders. H. MURADOUR AND G. AUBIN. *Compt. rend.* 192, 90-2 (1931), cf. C. A. 25, 818—The types of powder were a 20% centralite with a mixt. of nitrocotton 60, nitroglycerin 25, a 9% centralite powder made from the same mixt., a powder without centralite made from a mixt. of nitroglycerin 60, nitrocotton 50 and flake nitrocellulose powder. Various sizes of grains were used. Percent variation of $sp\,dt$ is independent of grain size except for very low d. of loading with 50% nitroglycerin powder. A graph is given, with abscissas the inverse of max. pressure, ordinates the inverse of $sp\,dt$, which shows that between 2500 and 3000 kg. max. pressure the inverse of $sp\,dt$ is practically a linear function of max. pressure.

E. M. SYMMES

Limits of inflammability of gases and vapors. H. F. COWARD AND G. W. JONES. *Bur. Mines, Bull.* 279, 114 pp. (1931), cf. C. A. 22, 2038—This is a thoroughly revised and enlarged edition of this *Bull.*

ALDEN H. EMERY

The ignition limits of hydrogen-air mixtures in the explosion bomb. FR. DINTRELFN. *Tech. Mech. u. Thermodynamik* 1, 429-34 (1930).—Values of ignition limits of H₂ air mixts. given in the literature vary from 0.041 to 0.10 and from 0.60 to 0.80 parts H₂ by vol. for the lower and upper limits, resp., this is mainly due to widely different app. used. In the tests made photography was used to det. the pressure time diagram, and the movement of the flame front was recorded by spark photography with the Lindner app. (*Forschungsab. Gebiete Ingenieurwesen* No. 326 (1930)). It was found that in addn. to the lower limit there was a combustion limit which is explained by the influence of convection. The proof thereof was obtained by variation of test conditions and Schlieren photographs. The characteristic reactions at the upper limit vary widely from those at the lower limit. An explanation is attempted by the use of the Schlieren photographs. F. M. SYMMES

Nitration of diethylene glycol. WM. H. RIXENRACH AND H. A. AARONSON. *Ind. Eng. Chem.* 23, 160-3 (1931).—Com. diethylene glycol contains usually about 5% ethylene glycol. The latter was removed by fractional distn. and freezing. Pure diethylene glycol has a *p* of -11.5° instead of -11.3° given in the literature. There were investigated the effects of the ratio of HNO₃ to H₂SO₄ in the mixed acid, the % H₂O in the mixed acid, the ratio of HNO₃ to diethylene glycol in nitration, the soly. of the dinitrate in spent acid, the effect of temp. in nitration, oxidation loss in nitration and purification losses. Diethylene glycol dinitrate is unstable in the presence of acid, is easily decomposed by heat and empirical heat tests are no more applicable than with nitroglycerin. Diethylene glycol dinitrate remained stable for 1 yr. at ambient temp. It is toxic like nitroglycerin. F. M. SYMMES

High brisance studies. III. Dynamite and Penthrinit in plastic and frozen states. ALFRED STETTBACHER. 2. *ges. Schuss-Sprengstoffe* 24, 8-10 (1931), cf. C. A. 25, 1674.—Nitroglycerin, nitroglycol or gelatin Penthrinit frozen at a temp. as low as -40° are more brisant than when in liquid or plastic form, as shown by tests on 7-mm. Fe plates. At -75° to -80° glycol dinitrate merely deflagrates, blasting gelatin explodes with greatly reduced brisance, pure nitroglycerin explodes with unaltered effect, while Penthrinit exhibits increased brisance. S. believes that these differences in behavior cannot be satisfactorily explained by differences in them or phys. properties. The tests made serve to emphasize the high brisance of the Penthrinit under all conditions. C. G. STONE

Firedamp, its characteristics and detection. COLBY MCLUCKIE. *Iron Coal Trades Rev.* 121, 800-1 (1930).—The amt. and rate of emission of gas from coal were detd. They vary widely in different seams, varying from a trace to 8.8 cu. ft. per ton of coal in 48 hrs. in the samples reported. Sampling of the gas by means of rubber pump and balloon is described. LESLIE B. BRAGE

The testing of blasting caps. A. HAID AND H. KOENEN. 2. *ges. Schuss-Sprengstoffe* 25, 393-8, 433-9, 463-8 (1930).—The general principle proposed by Grop and developed by Wohler (C. A. 21, 3744) of testing blasting caps against desensitized nitro compds. gives satisfactory results, provided test conditions are uniform. TNT of standard purity is screened through 0.5 mm. mesh with a duster screen of 0.15-mm. mesh. Talc was selected as the most desirable desensitizing agent, with specifications for screen test and setting test. The TNT and talc are intimately mixed in varying proportions, the talc content varying by 5% increments, and the mixts. are pressed at 1250 kg./cm² in blocks 25 mm. in diam. and 41-2 mm. high, with a cavity for insertion of the blasting cap. These blocks do not become insensitive on storage as do mixts. contg. oils. The blocks are detonated on Pb plates 3 cm. thick, and the diam. of the depression is measured after filing the rough edges. This diam. decreases with increasing talc content until it is only the diam. of the block when the latter fails to detonate. The initiating effect of a blasting cap is measured by the percentage to which the talc may be increased before failure results. Straight Hg(ONC)₂ caps failed at 25% talc compound TNT caps at 50% and tetryl caps at 65%. A series of tests was also made by the Cu cylinder method of Kast and the Trauzl Pb block test, which compared the effects of TNT charges desensitized with 30% talc when initiated with caps loaded to equal height with (A) Hg(ONC)₂ and KClO₄, 90/10, (B) Hg(ONC)₂, (C) TNT, (D) tetryl, (E) pentaerythritol tetranitrate. In C, D and E, priming charges of a mixt. of Pb trinitroresorcinate and PbN₆ were used. The order of brisance was found to increase from A to E. In every case, increased d. of charge in the caps gave higher brisance. The latter was shown to vary with the rate of detonation of the cap charge. Brisance is calcd. as $\Delta \times \text{rate of detonation} \times (T/273) \times V$. General confirmation of the above relations was also obtained in tests of the caps alone on 7-mm. Pb plates. Variations in the form of the base of the cap (flat or concave) do not affect the initiating

effect of the cap as they do the results of the Pb plate test. Both Cu and Al caps gave substantially the same results. In studying the effect of variations in granulation of the TNT used in the desensitized charges, it was found that granulations from 0.15 to 0.50-mm mesh showed equal sensitiveness (desensitization limit 10% tale), the same TNT pulverized to dust in a ball mill was more sensitive (50% tale), when it was fused and pulverized, its sensitiveness decreased (30-35% tale). Variations in pressures used in blocking the TNT tale mixts had a negligible effect on the results, with pressures from 750 to 2000 kg./cm.² the desensitization limit was the same (10% tale). Lower pressures gave a decrease in sensitiveness. An app. is described and illustrated for cutting open blasting caps in order to remove the charge intact for examination and analysis. C. G. STORM

Explosions occurring during the use of ethylene. FRANKLIN D. JOHNSON AND HUGH CANOT *Arch. Surg.* 22, 195-221 (1931)

Under-water explosions—"torpedo" effect. M. TONFUTTI / *ges. Schiess-Apparat* 25, 4 to 7 (1930). Tests conducted by T. lead to conclusions exactly opposite to those of Stettbacher (C. A. 24, 6018). Charges of 11 g TNT (d 1.1) and of 17 g Penthril 80/20 (d 1.1) in cylindrical metal cartridges 21 mm in diam were detonated on 6.5-mm Fe plates resting on an Fe cylinder enclosing a sealed air chamber. Tests were made in air and also with the entire app. under 0.5 m of water. Photographs of the Fe plates show that with both explosives the under water effect was appreciably greater than that obtained in air. C. G. STORM

Explosion in an acid-mixing plant. S. H. NEWMAN AND H. E. WATTS *British Home Office Report* 1931, 18 pp.—The explosion occurred in an acid mixing tank in the plant of Hickson and Partners, Ltd., Castleford, Yorkshire, Eng., on July 4, 1930. HNO_3 (97-98%) was being added to 2-3 tons of H_2SO_4 (approx. 80%) recovered from spent acid from the manu. of mono- and di-nitro compds of C_6H_5 and PhMe . Violent evolution of fumes was followed by flame and an explosion which caused much damage to the plant and the loss of 13 lives. Investigation showed that the storage tank from which the H_2SO_4 was drawn contained large quantities of a mixt. of nitro compds resulting from long accumulation. When HNO_3 was added to samples of these nitro compds mixed with H_2SO_4 , a marked rise in temp. occurred with a strong evolution of fumes. A mixt. of the nitro compds (2 vols.) with HNO_3 (3 vols.) was highly combustible. Of this mixt. 376 g was loaded in on 18 pounder steel artillery shell, provided with a booster of 50 g. tetryl, in which a no. 8 detonator was placed. On firing the filler, the charge detonated, causing fragmentation equal to that produced by pierce acid. It was concluded that the reaction of the HNO_3 with the nitro compds in the H_2SO_4 caused the explosion of the mixing tank. C. G. STORM

Nitro and amino compounds. Safe practices in production of explosives materials (AINSWORTH) 23. Starch explosion hazards reduced by safety measures (PICKER, BROWN) 12. Nitrocellulose from wood-pulp board (Brit. pat. 330,235) 23. Recovery of HNO_3 (Can. pat. 308,579) 18.

Explosives. CHARLES V. A. E. BARRON. Fr. 691,466, July 29, 1929. Explosives are coated with coal or vegetable tar, petroleum, etc., by dissolving the coating substance in a suitable solvent such as CCl_4 or mineral oils, mixing with the explosive and heating, with stirring, in an autoclave. The solvent is afterward recovered by distn. under vacuum.

Explosives. WALTER FAEDERICH. Fr. 694,057, April 16, 1930. Cyclic ketones or cyclic aces such as tetramethylcyclohexanone, tetramethylcyclopentanone or octomethylcyclohexanediol are converted into nitrates. The nitrates may be used alone or mixed with other explosives.

Explosives. OKAR MATTER. Ger. 513,933, May 27, 1928. Pure, readily explosive Pb azide is prepd. by treating Na azide with alk. earth soln. (e.g., $(\text{AcO})_2\text{Ba}$) and then with PbCO_3 .

Explosives. LEONARDO PARODI-DELFINO. Ger. 513,567, Oct. 30, 1925. Nitro powders, with or without a nitroglycerol content, are gelatinized and stabilized by addn. of phthalic or homologs. Thus, artillery explosive may consist of nitrocellulose, nitroglycerol and phthalic.

Nitroglycerin. A. SCHMIDT AND J. MEISSNER. Brit. 330,253, July 8, 1929. Sepn. of nitroglycerin from residual acid is effected by passing the crude liquid through a vessel having an inclined longitudinal axis so that the particles move generally parallel to this axis, the acid downwardly and the nitroglycerin upward. Various details of app. and operation are described.

"Gas" cartridge. BYRON C. GOSS (to Lake Erie Chem. Co.). U. S. 1,792,010 Feb 10. A compn contg gelatin and glycerol is used as the container material of cartridges such as those contg an asphyxiating or disabling chemical to be fired from a pistol.

Fire-producing compositions. FERDINAND RINGER. Austrian 120,172, June 15, 1920. A fire producing compn cast in stick or rod form comprises a combustible base giving only a little ash, e. g., cellulose nitrate or acetate, together with oxidizing agents and readily combustible substances, e. g., KClO_4 and S, with, if desired, additives for regulating the rate of combustion, e. g., $(\text{NH}_4)_2\text{C}_2\text{O}_4$. The proportions of the ingredients are selected so that the stick or rod can be ignited by friction, blown out and ignited again as required. Glass powder, Carborundum or the like may be included to increase the sensitiveness to friction. A suitable compn is cellulose nitrate 60-75, camphor 6-14, KClO_4 11-20, $\text{K}_2\text{Cr}_2\text{O}_7$ 2.5, $(\text{NH}_4)_2\text{C}_2\text{O}_4$ 2.5, S 2.5, glass meal 5-15 and ZnO 3-10%.

25-DYES AND TEXTILE CHEMISTRY

L. A. OLNEY

The chemistry of turkey-red dyeing. LITTLE R. PARKS. *J. Phys. Chem.* 35, 482-510 (1931).—The history of turkey red dyeing is reviewed and the chemistry of the various operations studied. A new short process is proposed which gives a color just as bright and fast as any produced by the older methods. H. W. LEAHY

Nitro and amino compounds. Safe practices in production of dyestuffs and explosives materials. CYRIL ARNSWORTHY. *Am. Dyestuff Repr.* 20, 125-7 (1931).—Prevention methods and remedies for poisoning by various nitro and amino compds. are given. ROBERT HUGHSON

Formic acid in the textile industry. K. L. SCHANBACHER. *Proc. Am. Assoc. Textile Chem. Colorists* 1931, 76-9, *Am. Dyestuff Repr.* 20, 144-7 (1931).—The use of HCOOH in the dye bath as a swelling agent, in cross dyeing and as an exhausting agent is discussed. MILTON HARRIS

Use and misuse of chlorine in textile bleaching. W. L. SAYRELL AND J. W. JVEY. *Am. Dyestuff Repr.* 20, 97-100 (1931).—Cl bleaches may be used from atmospheric temps. (cold) to 120°F in the form of strongly alk. to slightly acid hypochlorites. Bleaches are prepd. by dissolving Cl in caustic soda soln. NaHCO_3 is unsuitable and unnecessary as a constituent of bleach liquors. Stock solns are generally made at a concn. of approx. 20 g per l of available Cl. MILTON HARRIS

Bleaching with liquid chlorine or hydrogen peroxide 100 volume for desizing. C. H. DEDONIS. *Proc. Am. Assoc. Textile Chem. Colorists* 1931, 83-4; *Am. Dyestuff Repr.* 20, 151-2 (1931).—A discussion of formulas and their application in bleaching. MILTON HARRIS

The chemical analysis of silk cloth. C. SMITH. *Am. Dyestuff Repr.* 20, 118-22 (1931).—See C. A. 24, 3115. MILTON HARRIS

Change of some properties of sericin particles on the surface of the (silk) cocoon on drying. H. KANEKO, T. HAYASHI, S. CHINO AND M. MIYASAKA. *Bull. Sericult. Suk Ind. (Japan)* 2, 2-3 (1930).—During the drying of the cocoon the sericin particles on the surface undergo dehydration and their phys. properties change in such a way that the amt. of dyestuff absorbed is decreased and the surface tension, viscosity, n and turbidity and colloidal properties of the sericin soln. in water are altered. B. C. A.

Boiling-resistance tests of acetate silk. WALTER FERMAZIN. *Chem.-Ztg.* 55, 83-4 (1931).—Methods are described for testing the resistance to boiling of acetate silk yarn. Skeins are washed in ether, then tested by boiling for 1 hr. in distd. water and dried for 2 days at room temp. For testing breaking strength, elongation and elasticity, the boiled skeins are kept for 48 hrs. in 60% humidity, and 100 threads are then tested in the Schopper app. Elasticity values obtained by loading a weighted thread with an added weight equal to 50% of the breaking strength, and measuring the contraction on unloading, agree with the results of the other test. The depth of dyeing of skeins with standard dyes is compared with that of standard boiled skeins of varying resistance. The loss of gloss is detd. by Kempf's app., in which light is reflected from the threads onto a photographic film and the resultant brightening of the film is measured photometrically. I. H. ORELL

Determination of μ of cotton cloth and study of its relation to tensile strength. S. I. KOLSKY AND B. M. JONES. *Am. Dyestuff Repr.* 20, 133-8, 157-9 (1931).—Slight

decreases in p_n tender the cloth while increases tend to make it stronger. Ordinary bleached cloth was strengthened by treatment with dil alkali. MILTON HARRIS

The effect of cleansing agents on linen and cotton. EMIL J. SIMOLA. *Acta Chem. Fennica* 3, 89-91(1930)—One g per l. of $\text{MeCl}_2\text{SO}_2\text{NCINa}3\text{H}_2\text{O}$ has a little less weakening action on linen and cotton than has 0.5 g per l. of perborate bleaching agent because of the greater stability of the former on boiling. A mixt. of 5 g of soap with 3 g per l. of soda weakened cotton 15% in 25 washings and 22% in 50 washings. For linen the values were 25 and 62%, resp. For soda alone (0 g per l.) the values were 0 and 12%, resp. for cotton and 6 and 3%, resp. for linen. Ten g per l. of Marseille's soap, with a water content of 26%, weakened cotton 16% in 25 washings and 28% in 50 washings, with values of 26 and 46%, resp. for linen. Soaps prep'd from linseed oil, soy bean oil, Na and K stearates, oleates and palmitates showed that in water of 4 degrees hardness the greatest weakening of cloth occurred with soaps contg acids of high l no or high degree of unsatn. The cloth wt increased in all washings, being greatest when water of a high degree of hardness was used. The weakening of cloth when dist'd water, water of 4 degrees hardness and water of 25.4 degrees hardness were used were, resp., 5-15, 8-19 and 24-39%. For water of 4 degrees hardness the greatest weakening occurred with soaps contg Na salts of acids of high l no or high unsatn as in soy bean oil and linseed oil, being about 25% for cotton and 35% for linen in 25 washings. When soaps contg Na salts of sat'd acids as in coconut oil and tallow were used the weakening effect on cotton was 10% and on linen 15% in 25 washings. The drying of the cloth also greatly affects its strength, the weakening varying directly with the time necessary to dry. The cloth weakening was directly proportional to the amt. of Ca and Mg salts of unsat'd. acid ppt'd on the cloth fibers. These salts have some effect on the air oxidation of the cloth, and catalyze its autoxidation during drying, hence it is imperative that the cloth be dried as rapidly as possible. Ca and Mg salts must be prevented from forming, or, if this is not possible, they must be prevented from oxidation during drying of the cloth. Sat'd acids will cause a pptn on the cloth fibers, but the salts are not oxidized, and have less weakening action on the cloth. Sat'd acids prep'd from unsat'd acids were difficult to purify, and discolored the cloth on long standing. Substances may be added which keep Ca and Mg salts in emulsion form and prevent pptn on the cloth fibers. One cu m. of water of 12 degrees hardness requires 2 kg. of Na soap to soften, and results in a pptn of Ca and Mg salts, so the most effective method of preventing their formation is to soften the water preliminary to its use. Perborate soaps have about the same effect on cloth weakening as soaps contg a high percentage of unsat'd acids. S. A. K.

Newer preparation for emulsifying, wetting out and washing. A. SALMONY-KARSTEN. *Z. ges. Textilind.* 33, 677(1930)—"Fierhenol M Superior" is described as a dyeing oil for rayon and rayon mixts with silk and wool, "Fierhenol B T Special" is an anhydrous coned product for degreasing vegetable fibers; "Appret-Fierhenol," a finishing oil stable with MgSO_4 ; "Neo Fierhenol," a wetting out agent with unusual brightening and softening properties, "Visco-Fierhenol F" and "Visco Fierhenol A" are for brightening and dyeing cotton and rayons; "Enfullon II" and "Enfullon Extra" for fulling and washing of wools. All of these products are stable to hard water, easily sol., rinse well and are inexpensive. H. F. LEUBOLD

The photoelectric measurement of luster (DESBLOIS) 2. Experiments with N fertilizers on cotton (MANDRUGI) 15. Textile soap (HALBERSTAOT) 27. Purely aliphatic strepto-pentamethin dyes (KÖNIG, REGNER) 10. Constitution of colorless and colored triphenylmethane derivatives (LIFSCHITZ) 10. Protecting [textile materials] from fungi (Brit. pat. 330,244) 16.

Annuaire des soies et des soies artificielles, 1930. Paris: L'Édition textile 757 pp. F. 55. Reviewed in *Chimie & Industrie* 24, 1257(1930).

HEGEL, K. T.: *Textilchemische Erfindungen*. Lfg. 6. Wittenberg: A. Ziemsen Verlag. Paper, M. 10. Reviewed in *J. Soc. Dyers Colourists* 47, 40(1931). Cl. C. A. 24, 3377.

Dyes. SCOTTISH DYES, LTD. Ger. 512,821, May 12, 1927. See Brit. 278,030 (C. A. 22, 2667).

Dyes. I. G. FARBENIND A.-G. (Wilhelm Eckert, inventor). Ger. 513,000, Oct. 23, 1928. N-contg dyes are produced by nitrating naphthoylethylenediamine and

reducing the product. Thus 1,4,5,8-naphthoylenedibenzimidazole (obtained by condensing 1,4,5,8-naphthalenetetracarboxylic acid with *o*-phenylenediamine as described in Ger 430,632) is nitrated with HNO_3 and reduced with $\text{Na}_2\text{S}_2\text{O}_4$ to give a dye which colors cotton scarlet in a green vat. Other examples are given.

Dyes. I G FARBERINO A-G. Brit 337,021, June 21, 1929. Anthraquinone dyes for wool are made by condensing the Cu compd of a 1 hydroxy-4 haloanthraquinonesulfonic acid with an alkyl, aryl, cycloalkyl, aralkyl or acylamino compd, the amino residue entering the 4 position. Suitable starting materials include acetyl *p*-phenylenediamine, *m*-phenylenediamine and acetyl *m*-toluylenediamine.

Vat dyes. L CASPILLA & Co, Ges. Brit 330,425, Dec 10, 1928. See Ger 507,518 (C A 25, 823).

Dyes. SOC ANON FOUR L'IND CHIM A BALF. Brit 336,775, Oct 3, 1929. Isodibenzanthrone is chlorinated (1) with SbCl_5 in PhNO_2 to form a blue violet dye, (2) with sulfonyl chloride and I in PhNO_2 , or (3) with Cl and I in PhNO_2 , to form a similar dye, or (4) with I_2Cl_2 in trichlorobenzene, also forming a similar dye.

Dyes (dibenzanthrone derivatives). I B ANDERSON, R T THOMSON, J THOMAS and SCOTTISH DYES, LTD. Brit 330,268, May 9, 1929. By the alkali fusion of *Bz*: 1, *Bz*: 1'-dibenzanthronyls substituted by phenoxy, nitro or simple amino groups (but having the 2 position free) dyes are obtained producing reddish blue, greenish black or greenish gray shades on cotton. Several examples are given.

Azo dyes. I G FARBERINO A-G. Brit 336,938, July 10, 1929. Dyes are formed in substance or on the fiber (which may be regenerated cellulose or cellulose esters or ethers) by coupling a diazo, tetrazo or diazoazo compd free from sulfonic, hydroxy and carboxy groups with a 2',3'-hydrazynaphthyl-4-amino-1-alkoxymethylbenzene. Numerous examples of coupling components are given.

Azo dyes. I G FARBERINO A-G. Brit 335,893, April 30, 1929. A diamine of the general formula $\text{NH}_2\text{--R--R'--R--NH}_2$, in which R is an aromatic and R' a hydroaromatic residue (all of which may be substituted or not) is tetrazolized and coupled with 2 mol proportions of the same or different coupling components one of which may be a diazo compd or may be diazotized and coupled with another component. Numerous examples are given of dyes of different colors for dyeing wool, silk, etc.

Azo dyes. SOC ANON FOUR L'IND CHIM A BALF. Brit 335,890, June 25, 1929. Cotton is dyed olive green in the presence of Cr salts by the dye produced by coupling 2 sulfo-4-oxy- α -naphthophenazine with diazotized 4-nitro-2-amino-1-phenol-6-sulfonic acid, and instead of the latter compd there may be employed *o*-anisidine, *p*-nitroaniline, sulfamic acid or *p*-nitroaniline-*o*-sulfonic acid. Various other dyes also are described derived from components of similar type, and methods are given for producing azines used as intermediates.

Azo dyes. IMPERIAL CHEMICAL INDUSTRIES, LTD., and R BRIGHTMAN. Brit 336,350, Aug 15, 1929. Azo dyes for wool are formed by coupling tetrazotized *m,m'*-tolidine with one mol proportion of salicylic acid and one mol proportion of 2-naphthol, a 2-naphtholmonosulfonic acid, 2-naphthylamine-6-sulfonic acid or 2-methylamino-naphthalene-7-sulfonic acid. Cf C A 25, 699.

Azo dyes. IMPERIAL CHEMICAL INDUSTRIES, LTD. Fr 694,730, April 29, 1930. New azo dyes are prep'd by diazotizing a *p*-nitroamine or a *p*-acetamidoamine of the C_6H_4 series contg at least one alkoxy group in the mol, combining with a naphthol, naphtholsulfonic acid or a *N*-substituted 2,8-aminonaphtholsulfonic acid, reducing or hydrolyzing the product obtained, diazotizing again and coupling with a 1,8- or a 2,8-aminonaphtholmono- or disulfonic acid. Thus, 5-nitro-2-anisidine is diazotized and coupled with 2-phenylamino-8-naphthol-6-sulfonic acid. The product is reduced, diazotized and coupled with 1,8-aminonaphthol-2,4-disulfonic acid, giving a product which dyes viscose silk a deep blue. Other examples with a list of components and the colors obtained are given.

Azo dyes. IMPERIAL CHEMICAL INDUSTRIES, LTD. Fr 694,559, April 26, 1930. Azo dyes are prep'd by coupling with coupling components a tetrazotized diamine of the general formula $\text{NH}_2\text{C}_6\text{H}_3\text{NHNHC}_6\text{H}_3\text{NH}_2$, in which X is a chain of 2 or more CH_2 groups or a chain of CH_2 and CO groups, or by combining a diazotized amine of the general formula $\text{YC}_6\text{H}_4\text{NHNHC}_6\text{H}_3\text{NH}_2$, in which X has the same significance and Y is a NO_2 or acylamino group, with a constituent which contains no NH_2 group, reducing the NO_2 group or hydrolyzing the acylamino group, diazotizing and coupling with another constituent. The dyes give regular colors on regenerated cellulose silk.

Copper compounds of azo dyes. I G FARBERINO A-G. Brit 336,971, July 22, 1929. Substantive *o*-carboxyazo dyes contg Cu are formed by the action of a copperiz-

ing agent such as CuSO_4 on dyes obtained by coupling tetrazotized 4,4'-diaminodiphenyl 2,3'-dicarboxylic acid with 2 similar or dissimilar mol proportions of an amino-naphtholsulfonic acid or a nuclearly substitution product of such an acid (other than acids substituted in the amino group). The dye may also be prepd in the presence of a suitable coupling agent. The dyes thus formed produce mainly blue and gray shades on cotton and viscose. Several examples are given.

Monazo dyes. I. G. FARRINGTON, A-G. Brit 336,580, June 14, 1929. Monazo dyes which dye wool from an acid bath bright yellow to red to violet to brownish shades fast to washing and to mulling are formed by combining with a usual coupling component contg at least one sulfone group a diazotized aniline deriv. contg a hydrogenated benzene nucleus in *p*-position to the NH_2 group and which may also contain a halogen atom or an alkyl or alkoxy group in *o*-position to the NH_2 group. Numerous examples are given.

Disazo dyes. H. I. DU PONT DE NEMOURS & CO. Brit 336,646, July 12, 1929. Disazo dyes for cotton are made by coupling disazo compds. free from sulfo and carboxy groups with a *N*-aminobenzenylaminobenzoic *J*-acid, diazotizing the monazo product and coupling it with a sulfo- or carboxy-indole. Numerous examples are given.

Disazo dyes. J. R. GRAY, A-G. Brit 337,224, Nov 13, 1928. See Fr 681,851 (C. A. 24, 3507).

Disazo dyes. MELENIOR HORNIGER (to Chemische Fabrik vorm. Sanitex), U. S. 1,792,333, Feb. 10. By coupling one mol of a tetrazotized phenyl with two mols of 3'-amino-4'-methyl-5'-sulfo-1-phenyl-5-pyrazolones or by combination of the intermediate products of one mol of a tetrazotized phenyl and one mol of an aryl-*o*-hydroxybenzoic acid with one mol of a 3'-amino-4'-methyl-5'-sulfo-1-phenyl-5-pyrazolone or by coupling the intermediate product of one mol of a tetrazotized phenyl and one mol of a 3'-amino-4'-methyl-5'-sulfo-1-phenyl-5-pyrazolone with one mol of an azo component substituent aryl dyes are obtained, which when diazotized on the fiber can be developed with unfluorinated azo components (for instance *p*-naphthol, *m*-diamines, phenylmethyl-pyrazolones) to shades varying from orange to red, heliotrope and dark brown, which are fast to washing and can be discharged with formaldehyde-hyposulfite to pure white effects. The direct dyeings can also be fixed with formaldehyde or by an after treatment with *p*-nitrochlorobenzene. The same diazotizable azo dyes can be obtained by substituting, for the 3'-amino-4'-methyl-5'-sulfo-1-phenyl-5-pyrazolones, the corresponding 3'-nitro-4'-methyl-5'-sulfo-1-phenyl-5-pyrazolones, and by reducing in the nitro dyes thus obtained the nitro group to the amino group by treatment with Na_2S . Several examples with details of procedure are given.

Disazo dyes for acetylcellulose. HERRON DYEING CORP., LTD., JAMES BADDILLY and JAMES HILL, Ger. 513,763, Nov. 9, 1928. See Brit 270,428 (U. S. A. 22, 1692).

Vat dyes. I. G. FARRINGTON, A-G. (Max A. Kuhn, Karl Koberle and Erich Berthold, inventors), Ger. 516,313, Apr. 25, 1929. Green blue vat dyes contg halogen are prepd by treating benzanthraquinonecarthronic or its derivs. with HClO_4 in the presence of a catalyst, *e. g.*, Hg or I , and in the presence or absence of Br or other brominating agent. The products give greener shades than the dyes described in Ger. 492,274 (C. A. 24, 2011). Examples are given.

Vat dyes. I. G. FARRINGTON, A-G. (Heinrich Neresheimet, inventor), Ger. 516,281, Aug. 15, 1929. Addn. to 513,918 (C. A. 23, 1390). Blue to green vat dyes similar to some of those obtainable as described in Ger. 513,918, are prepd by the action of acid condensing agents on *o*-anthraquinonyl 3-amino-1,2-benzanthraquinones contg at least one acylanino group, which may be split off or substituted after the condensation. Thus, the condensation product from 3-chloro-1,2-benzanthraquinone and 1-amino-5-benzoylanilinoanthraquinone may be treated with HClO_4 . Examples are given.

Vat dyes. I. G. FARRINGTON, A-G. Fr. 691,152, April 18, 1930. The 4 H atoms of the C_{11} ring in anthraquinone-2,1 acridone are substituted by halogen atoms. The dyes are prepd after the process of Ger. 267,211, *e. g.*, by condensing the benzylic ester of 1-chloroanthraquinone-2-carboxylic acid with 2,3,4,5-tetrahaloaniline and a final joining of the acridone ring, or by introducing other halogen atoms into anthraquinoneacridones, the C_{11} ring of which is substituted by less than 4 halogen atoms, or by substituting with H the halogen contained (in addn. to 4 atoms of halogen of the C_{11} ring) in another part of the mol of the halanthraquinoneacridones. Examples are given.

Vat dyes. I. G. FARRINGTON, A-G. Fr. 693,997, April 13, 1930. The mists of isomeric 1,4,5,8-naphthoquinediarylimbdiaroles obtained by the process of Fr. 693,813

reducing the product. Thus, 1,4,5,8-naphthoylenedibenzimidazole (obtained by condensing 1,4,5,8-naphthalenetetracarboxylic acid with *o*-phenylenediamine as described in Ger 430,632) is nitrated with HNO_3 and reduced with $\text{Na}_2\text{S}_2\text{O}_4$ to give a dye which colors cotton scarlet in a green vat. Other examples are given.

Dyes. I G FARBEIND A-G Brit 337,021, June 21, 1929 Anthraquinone dyes for wool are made by condensing the Cu compd of a 1-hydroxy-4-haloanthraquinonesulfonic acid with an alkyl, aryl, cycloalkyl, aralkyl, or acylamino compd, the amino residue entering the 4 position. Suitable starting materials include acetyl-*p*-phenylenediamine, *m*-phenylenediamine and acetyl-*m*-tolylenediamine.

Vat dyes. L CASSELLA & Co, Ges Brit 336,495, Dec 10, 1928 See Ger 507,558 (C A 25, 823)

Dyes. SOC ANON FOUR L'IND CHIM A BALE Brit 336,775, Oct 3, 1929 1-isodibenzanthrone is chlorinated (1) with SbCl_5 in PhNO_2 to form a blue violet dye, (2) with sulfuric chloride and *I* in PhNO_2 or (3) with Cl_2 and *I* in PhNO_2 , to form a similar dye, or (4) with FeCl_3 in trichlorobenzene, also forming a similar dye.

Dyes (dibenzanthrone derivatives). I B ANDERSSON, R. F. THOMSON, J THOMAS and SCORRIN DYES, LTD Brit 336,268, May 9, 1929 By the alkali fusion of *Bz*-1,3'-1'-dibenzanthronyls substituted by phenoxy, nitro or simple amino groups (but having the 2 position free) dyes are obtained producing reddish blue, greenish black or greenish gray shades on cotton. Several examples are given.

Azo dyes. I G FARBEIND A-G Brit 336,938, July 10, 1929 Dyes are formed in substance or on the fiber (which may be regenerated cellulose or cellulose esters or ethers) by coupling a diazo, tetrazo or diazoazo compd free from sulfonic, hydroxy and carboxy groups with a 2'-3'-hydroxynaphthoyl-4-amino-1-alkoxymethylbenzene. Numerous examples of coupling components are given.

Azo dyes. I G FARBEIND A-G Brit 335,893, April 30, 1929 A diamine of the general formula $\text{NH}_2\text{--R--R'--R--NH}_2$, in which R is an aromatic and R' a hydroaromatic residue (all of which may be substituted or not) is tetrazotized and coupled with 2 mol proportions of the same or different coupling components one of which may be a diazo compd or may be diazotized and coupled with another component. Numerous examples are given of dyes of different colors for dyeing wool, silk, etc.

Azo dyes. SOC ANON FOUR L'IND CHIM A BALE Brit 335,896, June 25, 1929 Cotton is dyed olive green in the presence of Cr salts by the dye produced by coupling 2-sulfo-4-oxo- α,β -naphthophenazine with diazotized 4-nitro-2-amino-1-phenol-6-sulfonic acid, and instead of the latter compd there may be employed *o*-anisidine, *p*-nitroaniline, sulfanilic acid or *p*-nitroaniline-*o*-sulfonic acid. Various other dyes also are described derived from components of similar type, and methods are given for producing azines used as intermediates.

Azo dyes. IMPERIAL CHEMICAL INDUSTRIES, LTD., and R BRIGHTMAN Brit 336,350, Aug 15, 1929 Azo dyes for wool are formed by coupling tetrazotized *m,m'*-tolidine with one mol proportion of sabcylic acid and one mol proportion of 2-naphthol, a 2-naphtholmonosulfonic acid, 2-naphthylamine-6-sulfonic acid or 2-methylaminonaphthalene-7-sulfonic acid. Cf C A 25, 599

Azo dyes. IMPERIAL CHEMICAL INDUSTRIES, LTD Fr 694,730, April 29, 1930 New azo dyes are prepd. by diazotizing a *p*-nitroamine or a *p*-acetamidoamine of the C_6H_4 series contg at least one alkoxy group in the mol, combining with a naphthol, naphtholsulfonic acid or a *N*-substituted 2,8-aminonaphtholsulfonic acid, reducing or hydrolyzing the product obtained, diazotizing again and coupling with a 1,8- or a 2,8-aminonaphtholmono- or disulfonic acid. Thus, 5-nitro-2-anisidine is diazotized and coupled with 2-phenylamino-8-naphthol-6-sulfonic acid. The product is reduced, diazotized and coupled with 1,8-aminonaphthol-2,4-disulfonic acid, giving a product which dyes viscose silk a deep blue. Other examples with a list of components and the colors obtained are given.

Azo dyes. IMPERIAL CHEMICAL INDUSTRIES, LTD Fr 694,559, April 26, 1930 Azo dyes are prepd. by coupling with coupling components a tetrazotized diamine of the general formula $\text{NH}_2\text{C}_6\text{H}_4\text{NHNHC}_6\text{H}_4\text{NH}_2$, in which X is a chain of 2 or more CH_2 groups or a chain of CH_2 and CO groups, or by combining a diazotized amine of the general formula $\text{YC}_6\text{H}_4\text{NHNHC}_6\text{H}_4\text{NH}_2$, in which X has the same significance and Y is a NO_2 or acylamino group, with a constituent which contains no NH_2 group, reducing the NO_2 group or hydrolyzing the acylamino group, diazotizing and coupling with another constituent. The dyes give regular colors on regenerated cellulose silk.

Copper compounds of azo dyes. I G FARBEIND A-G Brit 336,971, July 22, 1929 Substantive *o*-carboxyazo dyes contg Cu are formed by the action of a copperiz-

ing agent such as CuSO_4 on dyes obtained by coupling tetrazotized 4,4'-diaminodiphenyl-3,3'-dicarboxylic acid with 2 similar or dissimilar mol proportions of an aminonaphtholsulfonic acid or a nuclearly substitution product of such an acid (other than acids substituted in the amino group). The dye may also be prepd in the presence of a suitable coppering agent. The dyes thus formed produce mainly blue and gray shades on cotton and viscose. Several examples are given.

Monoazo dyes. I G FARBEIND A-G Brit 336,580, June 14, 1929. Monoazo dyes which dye wool from an acid bath bright yellow to red to violet to brownish shades fast to washing and to milling are formed by combining with a usual coupling component contg at least one sulfone group a diazotized aniline deriv contg a hydrogenated benzene nucleus in *p*-position to the NH_2 group and which may also contain a halogen atom or an alkyl or alkoxy group in *o*-position to the NH_2 group. Numerous examples are given.

Disazo dyes. E I DU PONT DE NEMOURS & Co Brit 336,646, July 12, 1929. Disazo dyes for cotton are made by coupling diazo compds free from sulfo and carboxy groups with a *N* aminobenzoylamino benzoyl-J-acid, diazotizing the monoazo product and coupling it with a sulfo- or carboxy-indole. Numerous examples are given.

Disazo dyes. J. R. GEIGY A-G Brit 337,224, Nov 15, 1928. See Fr. 684,551 (C A 24, 5507).

Disazo dyes. MELCHIOR BOEWIGER (to Chemische Fabrik vorm. Sandor), U S 1,792,355, Feb 10. By coupling one mol of a tetrazodiarlyl with two mols. of 3'-amino-4'-methyl-5'-sulfo-1-phenyl-5-pyrazolones or by combination of the intermediate products of one mol of a tetrazodiarlyl and one mol of an aryl-*o*-hydroxycarboxylic acid with one mol of a 3'-amino-4'-methyl-5'-sulfo-1-phenyl 5-pyrazolone or by coupling the intermediate product of one mol. of a tetrazodiarlyl and one mol of a 3'-amino-4'-methyl-5'-sulfo-1-phenyl-5-pyrazolone with one mol. of an azo component substantive azo dyes are obtained, which when diazotized on the fiber can be developed with unsulfonated azo components (for instance β naphthol, *m*-diamines, phenylmethyl pyrazolone) to shades varying from orange to red, heliotrope and dark brown, which are fast to washing and can be discharged with formaldehyde-by-sulfite to pure white effects. The direct dyeings can also be fixed with formaldehyde or by an after treatment with *p*-nitrodiazobenzene. The same diazotizable azo dyes can be obtained by substituting, for the 3'-amino-4'-methyl-5'-sulfo-1-phenyl-5-pyrazolones, the corresponding 3'-nitro-4'-methyl-5'-sulfo-1-phenyl-5-pyrazolones, and by reducing in the nitro dyes thus obtained the nitro group to the amino group by treatment with NaS . Several examples with details of procedure are given.

Diazo dyes for acetylcellulose. BRITISH DYEING CO., LTD., JAMES BADDILEY and JAMES HILL. Ger 513,763, Nov 9, 1926. See Brit. 270,428 (C. A. 22, 1692).

Vat dyes. I. G. FARBEIND A-G (Max A. Kunz, Karl Köberle and Erich Berthold, inventors) Ger. 516,313, Apr. 28, 1929. Green blue vat dyes contg. halogen are prepd. by treating benzanthronepyrazoleanthrone or its deriva. with HClSO_4 in the presence of a catalyst, *e.g.* Hg or I, and in the presence or absence of Br or other brominating agent. The products give greener shades than the dyes described in Ger 492,274 (C. A. 24, 2611). Examples are given.

Vat dyes. I. G. FARBEIND A-G (Heinrich Neresheimer, inventor). Ger. 516,784, Aug 18, 1929. Addn. to 513,046 (C A 23, 1396). Blue to green vat dyes similar to some of those obtainable as described in Ger 513 046, are prepd by the action of acid condensing agents on *o*-anthraquinonyl-3-amino-1,2-benzanthraquinones contg at least one acylamino group, which may be split off or substituted after the condensation. Thus, the condensation product from 3-chloro-1,2-benzanthraquinone and 1-amino-5-benzoylaminoanthraquinone may be treated with HClSO_4 . Examples are given.

Vat dyes. I. G. FARBEIND A-G Fr 694,152, April 18, 1930. The 4 H atoms of the C_6H_4 ring in anthraquinone-2,1 acridone are substituted by halogen atoms. The dyes are prepd. after the process of Ger 267,211, *e.g.* by condensing the benzylic ester of 1-chloroanthraquinone-2-carboxylic acid with 2,3,4,5-tetrahaloaniline and a final joining of the acridone ring, or by introducing other halogen atoms into anthraquinoneacridones, the C_6H_4 ring of which is substituted by less than 4 halogen atoms, or by substituting with H the halogen contained (in addn to 4 atoms of halogen of the C_6H_4 ring) in another part of the mol. of the haloanthraquinoneacridones. Examples are given.

Vat dyes. I. G. FARBEIND A-G. Fr. 693,997, April 15, 1930. The mixts. of isomeric 1,4,5,8-naphthoylenediazarylmudazoles obtained by the process of Fr. 600,843

and its addns are sep'd into their 2 constituents. The sep'n is effected by treatment with conc'd H_2SO_4 or by transforming them in the presence of an appropriate diluent into their salts and sep'g these. Examples and formulas of the products obtained are given.

Vat dyes. I G FARBERND A-G. Fr. 693,910, June 14, 1929. The condensation products obtained from *N*-dihydro-1,2',2,1'-anthraquinonazines by means of CH_3O or substances yielding CH_3O (cf. Fr. 346,398) are submitted to an oxidation or dehydrogenation. Examples are given.

Vat dyes of the thiondigo series. KARL THIAS, THEODOR MEISSNER and HANS HEYNA (to General Aniline Works). U. S. 1,792,130, Feb. 10, 1929. 4,7,4',7'-Tetramethyl-5,5'-dibromondigo and other similar halogen comp'ds. dye cotton blue from an olive-green vat and are made by transforming amino comp'ds. of the general formula

$CX-CMe-CY-C(NH_2)-CMe-CH$ wherein X stands for halogen and Y stands for H or the $COOH$ group, into indigo dyes according to the process generally used for the prep'n. of indigo dyes from aromatic amino comp'ds with a free *o*-position, or from *o*-aminocarboxylic acids. Several examples with details of procedure are given.

Vat dyes or intermediates. I G FARBERND A-G (Walter Mieg and Robert Berliner, inventors). Ger. 516,398, Dec. 9, 1928. Dianthraquinonylamine deriva. contg. a thioxanthone ring and either one acylimino group or an acridone ring is treated with H_2SO_4 at a low temp. Oxidation of the products may be necessary. Thus, the condensation product from 4-aminoanthraquinonethioxanthone and 1-benzoylamino-5-chloroanthraquinone may be treated with H_2SO_4 at 20° and the product poured into water contg. a little $NaNO_2$. Other examples are given also. The products are carbazole deriva. Cl. C. A. 25, 824.

Dyes and intermediates. W. SMITH, J. THOMAS and SCOTTISH DYES, LTD. Brit. 336,982, April 16, 1929. 1,1'-Dianthraquinonyl deriva. contg. Cl, Br or sulfonic acid groups in the 2,2'-positions are prep'd by condensing anthraquinone deriva. contg. similar groups in the 2 position and a halogen atom or diazo group in the 1-position in the presence of Cu. There may be employed as starting materials comp'ds. such as 1,2-dichloroanthraquinone, 1-amino-2-bromoanthraquinone or 1-amino-2-sulfanthraquinone. The 2,2'-substituents in the products are easily replaceable by other substituents, e. g., treatment with NH_3 forms 2,2'-diamino-1,1'-dianthraquinonyl which is converted to flavanthrone.

Dyes and intermediates. I G FARBERND A-G (Max A. Kunz, Karl Köberle and Gert Kochendörfer, inventors). Ger. 516,399, Feb. 16, 1929. See Fr. 689,969 (C. A. 25, 1100).

Dyes and intermediates. I G FARBERND A-G. Fr. 694,853, May 2, 1930. Condensation products, some of which are dyes and intermediates, are prep'd. by treating with condensing agents the products resulting from the reaction of aromatic *o*-aminocarboxylic acids, their deriva. or substitution products with negatively substituted heterocyclic comp'ds. or with negatively substituted isocyclic comp'ds. contg. at least 4 rings, or the corresponding products resulting from the reaction of heterocyclic or isocyclic *N* comp'ds. contg. at least 4 rings and having at least one atom of H capable of being substituted fixed to the N, with aromatic carboxylic acids, negatively substituted in the *o*-position to the $COOH$ group, or their deriva. or substitution products. Thus, the condensation product of dibromo-3,4,8,9-dibenzopyrene-5,10-quinone with 2 mols. of the Et ester of 1-aminoanthraquinone 2-carboxylic acid is dissolved in conc'd. H_2SO_4 and heated to 110-120° for some time, the oxonium salt of the product being ppt'd by water. On hydrolysis a product is obtained which dyes cotton in brown-red shades from the vat. Other examples are given.

Dye intermediates. I G FARBERND A-G. Brit. 336,394, Sept. 27, 1929. *p*-Arylamino-phenolcarboxylic acids are made by treating a dry alkali salt of a *p*-aryl aminophenol contg. a free *o* position to the OH group and no sulfonic acid groups with CO_2 with use of heat and pressure. There may be used as starting materials 4-hydroxydiphenylamine, 4-hydroxy-2'-methyldiphenylamine, 4-hydroxy-3'-methyldiphenylamine, 4-hydroxy-4'-methyldiphenylamine, 4-hydroxy-3-methyldiphenylamine (obtained from toluhydroquinone and aniline), 4-hydroxy-3,4'-dimethyldiphenylamine (obtained from toluhydroquinone and *p*-toluidine), and 4-hydroxy-3'-methoxydiphenylamine.

Dye intermediates. I G FARBERND A-G. Brit. 336,428, Oct. 14, 1929. Arylides of *m*-arylamino-phenolcarboxylic acids such as those obtained according to the process of Brit. 333,783 (C. A. 25, 603) are made by reaction of the carboxylic group

(with or without previous acylating of the N atom) with an arylamine in the presence of an acid-condensing agent such as PCl_5 and a diluent such as xylene, toluene, PhCl or dimethylaniline, or, the carboxylic acid may be first converted to the acid chloride by a reagent such as thionyl chloride and then condensed with an arylamine in the presence of a diluent. Numerous examples are given.

Dye intermediates. I. G. FARBERND A-G. Brit. 336,800, Oct. 18, 1929. Dihydroxyanilides of naphthalenedicarboxylic acids, capable of being coupled with a diazo compd., are made by condensing a naphthalenedicarboxylic acid (other than a *para* or an *ortho* acid) in the presence of PCl_5 (or by using a naphthalenedicarboxylic acid chloride directly) with a hydroxyarylamine which has a free *o*- or *p*-position relative to the OH group such as 3- or 4-aminophenol, 1 methyl 2-amino-4-hydroxybenzene, 1 methyl 2-amino-5-hydroxybenzene or the like. The alkali salts of the products are sol. in water and have an affinity for vegetable fibers. Naphthalene-1,4- and 1,5-dicarboxylic acid chlorides (which may be used as initial materials) are made by treating the corresponding acids with PCl_5 .

Dye intermediates. I. G. FARBERND A-G. Brit. 337,047, July 26, 1929. Amino-hydroxyanthraquinones and their substitution products are obtained by condensing an acyl- or aroylaminophenol, or a homolog or substitution product, with phthalic anhydride, or a substitution product, in the presence of a substance such as AlCl_3 , saponifying, and ring closing when necessary in the usual manner. Examples are given of the production of 2,3-aminohydroxyanthraquinone from 2,2'-dimethoxydiphenylurea, 1-amino-4-hydroxyanthraquinone from 1-benzoylamino-4-phenol and of several other similar reactions.

Dye intermediate (1-chloromethylnaphthalene). I. G. FARBERND A-G. Brit. 337,289, Jan. 12, 1929. Cryst. naphthalene in aq. suspension is treated with non-polymerized formaldehyde and concd. HCl (suitably with gentle heating).

Intermediates for azo dyes. I. G. FARBERND A-G. Ger. 516,156, Feb. 8, 1929. Nitroaryl arylides of the formula $\text{HORNHCOR}'\text{NO}_2$, in which R and R' are the same or different aryl radicals of the C_6H_5 series, and in which the NO_2 group is in the *m*- or *p*-position to the CO group, while the *o*- or *p*-position to the OH group is unsubstituted, are converted by treatment with alk. reducing agents into new compds. having a great affinity for cotton. The products are believed to be azoxy or azo compds. Examples are given.

Dyes of the anthraquinone series. FRANZ ACKERMANN and CARL APOTHEKER (to Soc. anon. pour l'ind. chim. à Bâle). U. S. 1,792,343, Feb. 10. Reaction products such as those formed by heating, at 150-160°, aromatic amines such as PhNH_2 or *p*-aminidine with compds. such as (1,5- or 1,8)-diamino-(4,8- or 4,5)-dinitroanthraquinone, are subjected to careful reduction (suitably with Na sulfide or hydrosulfide), and there are thus obtained dyes sol. in various org. solvents and which are suitable for dyeing cellulose esters and ethers or products such as nitrocellulose lacquers blue to blue-green shades. Several examples with details of procedure are given.

Fuchsin dyes. HEIDELBERGER GELATINE-FABRIK STROESS & Co (Albert Steigmann, inventor). Ger. 513,610, Dec. 14, 1928. Fuchsin dyes are produced by oxidizing and nitrating fuchsin in the presence of sulfuring agents and optionally reducing the product. Thus, fuchsin is mixed with $\text{Na}_2\text{S}_2\text{O}_4$ or Na_2SO_3 and treated with HNO_3 and H_2SO_4 . The Na and NH_4 salts of the dye color wool yellow in a weak AcOH bath. Another example is given. Cf. C. A. 24, 4639.

Sulfur dyes. I. G. FARBERND A-G. Brit. 336,061, Sept. 2, 1929. Sulfur dyes are made by sulfuring a 2-(4'-hydroxyarylamino)-6-arylammonaphthalene (suitably with a polysulfide of high S content and in the presence of an inert org. solvent such as an aliphatic alc. of high b. p., a cyclic alc. or glycerol). The dyes produce bluish to green shades on cotton from a Na sulfide bath or hyposulfite vat, and greener shades are obtained by sulfuring in the presence of a Cu salt such as the sulfide or cyanide. Several examples are given.

Sulfuretted dyes. I. G. FARBERND A-G. Brit. 337,061, July 27, 1929. The S dyes obtainable, as described in Brit. 325,519 (C. A. 24, 4169), by heating halogenated dinaphthylene dioxides with S, with or without a high boiling diluent, or by submitting them to polysulfide fusion, are treated either in substance or on the fiber with alkylating or aralkylating agents as described in Brit. 317,775 (C. A. 24, 2306). Several examples are given.

Wool dyes. I. G. FARBERND A-G. Fr. 694,088, April 17, 1930. One mol. of a secondary amine of the C_6H_5 series contg. 2 primary NH_2 groups in position *o*- with respect to the secondary amine group and at least one SO_3H group is heated with one

mol of the $C_{12}H_8$ series contg 2 NO_2 groups in *o*-position and a NO_2 group in *p*-position with respect to an exchangeable atom or atomic group. Thus, 2,6-diaminodiphenylamine-4 sulfonic acid is dissolved in water and the neutralized soln is condensed at 70-80° with 2,4,6-trinitroanisole in the presence of Na_2CO_3 . The product is pptd. by $NaCl$ and dyes wool in yellowish brown shades. Other examples are given.

Anthanthrone derivatives. I G FARBENIND A-G (Max A. Kunz, Karl Köberle and Erich Berthold, inventors) Ger. 516,312, Aug 10, 1928. See Fr. 678,055, (C. A. 24, 3248).

Dibenzopyrenequinones. GEORG KRANTZLEIN and HEINRICH VOLLMANN (to General Aniline Works) U S 1,792,169, Feb 10. Vat dyes are formed by the reaction (with introduction of O into the reaction mixt) at temps. of 110-50° of α -benzoyl naphthalene or other suitable α -arylnaphthalenes with benzoyl chloride or other suitable compd. of the general formula $aryl-CO-X$ (in which X stands for a halogen atom or for the group $-O-CO-aryl$, as in aromatic carboxylic acid chlorides or carboxylic anhydrides), in the presence of $AlCl_3$. The reaction may require 20-60 hrs. 1 examples and details of procedure are given.

Diphthaloylnaphthalene. I G FARBENIND A-G Fr. 694,890, May 2, 1930. 1,2,5,6-Diphthaloylnaphthalene, m. 410°, and its substitution products are prepd by treating 2',2'-dicyano-1,5-dibenzoylnaphthalene with an acid-condensing agent. If desired the cyano groups may first be changed to $COOH$ groups. With sufficiently moderate conditions 2'-carboxybenzoylnaphthanthraquinone may be obtained as an intermediate product. The prepn. of 1,2,5,6-(2',2'-dichlorodiphthaloylnaphthalene), m. above 400°, is also described. The products are intermediates for dyes.

Dyeing with vat dyes. I G FARBENIND A-G (Ludwig Löchner, inventor). Ger 513,536, Sept 12, 1928. Polychromatic effects are produced by using ester salts of vat dyes. The goods are impregnated with an alk. soln. of 2,3-hydroxynaphthoic acid arylides, and printed out successively with aromatic diazo compds. of vat dyes, with addn. of $NaNO_2$. The vat dye is developed in the dried goods by acid and the rolled goods are then scoured and treated with a hot dil alk. soln. A small quantity of reducing agent such as CH_3O , glucose, etc., may be added to the vat dye printing paste. In an example, the goods are soaked in a soln. contg 2,3-hydroxynaphthoic acid-toluidine Turkey red oil and $NaOH$, and dried. For red shades, the goods are printed out with a paste contg the diazo soln. from 4-chloro-2-toluidine- HCl , water, $NaNO_2$, $AcONa$ and starch. For blue shades the material is printed out with a soln. contg ester salts of tetrabromoindigo, $NaNO_2$, water and starch. Several other examples are given.

Dyeing or treating threads. HENRY DREYFUS. Fr 694,880, May 2, 1930. Liquids such as size or dyes are applied to threads while they pass below the edge of an annular or cylindrical member serving to guide the thread to a bobbin or cake of thread.

Dyeing cotton. ORANIENBURGER CHEMISCHE FABRIK A-G (Kurt Lindner, Rudolf Kuhn and Max Trapke, inventors) Ger 513,843, Jan 11, 1924. Dyes, especially naphthol dyes are applied directly to cotton, the process being carried out in the presence of hydrogenated phenols and nonsulfonated soap in the proportions 1 or less 4. Thus, the azo dye from diaminidine and 1-amino-8-naphthol-4-sulfonic acid (Chicago blue B) is used directly to color cotton blue in the presence of coconut oil, oleic acid and cresol. Other examples are given.

Dyeing wool yarn. FIRMA H. KRANTZ. Ger 513,682, April 24, 1927. Details are given of app. for receiving the spools and supplying the dye liquor.

Dyeing artificial silk. IMPERIAL CHEMICAL INDUSTRIES, LTD. Fr 694,856, May 2, 1930. Textiles of regenerated cellulose are dyed regular blue shades by applying to the textile a diazo dye obtained by coupling a diazotized aminoazo compd. with a 1,8- or 2,8-aminonaphthol mono- or -di-sulfonic acid, the aminoazo compd. being obtained by coupling with a naphthol or naphtholsulfonic acid or with a *N*-substituted deriv. of 2,8-aminonaphtholsulfonic acid, either diazotized *p*-nitroaniline (or a substituted deriv. excepting alkoxy derivs.) or diazotized *p*-aminoacetanilide (or a substituted deriv. excepting alkoxy derivs.), the coupling being followed by reduction of the NO_2 group in one case or hydrolysis of the acetylamino group in the other. Examples and a list of components with the colors obtained are given. Fr 694,857 describes a process for obtaining regular dyes on textiles of regenerated cellulose by applying to the textile a monoazo dye obtained by diazotizing an oxamide acid of the benzidine series and coupling with a naphtholsulfonic acid, a sulfonic acid of naphthylamine, 1,8-aminonaphtholsulfonic acid, 1,8-dihydroxynaphthalenesulfonic acid, *m*-dihydroxybenzene deriv., *m*-diaminobenzene deriv. or *m*-aminohydroxybenzene deriv. Examples and a list of components with the colors obtained are given. Cf. C. A. 25, 1685.

Dyeing cellulose esters or ethers. I G FARBEINING A-G (Wilhelm Eckert, Carl E Muller and Walter Gmelin, inventors) Ger. 516,982, Feb 2, 1929. Addn to 515,029 (C A 25, 1685) The method of Ger 515,029 is modified by using naphthalimide-4 sulfamic acid or the corresponding derivs or their salts instead of 4 aminonaphthalimide or the derivs mentioned Examples are given

Dyeing skins, furs, feathers, etc. I G FARBEINING A-G (Karl Marx and Karl Bittner, inventors) Ger 513,525, Apr 14, 1927 Use is made of a soln contg an oxidizing agent and a salt of the base $p\text{-RC}_6\text{H}_4\text{NH}_2$, where $R=\text{H}$, Cl , OCH_3 , or halogen The material may be mordanted In the example, skins are treated with an aq soln of H_2O_2 and 4-hydroxyethylaminodiphenylamine-HCl to produce a yellowish tint when mordanted with Fe alum A blue-violet color is given by 4 hydroxyethylamino-4'-methoxydiphenylamine-HCl

Skein-dyeing apparatus. J SCHUMPF Brit 336,174, Dec 27, 1928 Structural features

Apparatus for dyeing width fabrics. O OBERMEYER Brit 336,724, Aug 26, 1929 Structural features

Apparatus for dyeing cloth webs. WM E CAFFRAY (to Nashua Mfg Co) U. S 1,791,954, Feb 10 Various structural details are described

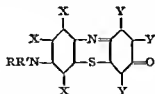
Apparatus for dyeing ribbons, etc. H A BLUNDELL and BLUNDELL BROS (LUTON), LTD Brit 336,686, July 23, 1929 Structural features

Extracting dyes from dyed artificial materials. WM WHITEHEAD (to Celanese Corp of America) U S 1,791,990, Feb 10 Dyed materials formed from artificial org cellulose derivs such as cellulose acetate are treated with a soln. such as $\text{C}_2\text{H}_5\text{Cl}$ together with about 2% of acetone, which acts as a swelling agent for the cellulose deriv material and a solvent for dyes present

Reduction products of dyes. I. G FARBEINING A-G Fr 693,904, April 14, 1930 Stable reduction compds which are difficultly sol in aq alkalies are prep'd by submitting to an alk. reduction the symmetrical or asymmetrical dyes of the thiondigo series which are substituted in their mol. at least once, by a halogen alkyl or alkoxy group, the alkali being present in amt. less than that used practically for the prep'n. of the vat dyes in question Several examples are given

Printing textiles. I G FARBEINING A-G Fr 693,996, April 15, 1930 Textiles are printed with stable reduction products of thiondigo dyes prep'd by the process described in Fr 693,904 (preceding abstr) and then treated in the usual manner. Examples are given. Cf C A. 25, 1391

Textile printing. I G. FARBEINING A-G Fr. 694,643, April 28, 1930 Fast printings are obtained on vegetable fibers by direct printing in the usual way with dyes of the formula:



in which R' is an aryl ring which may be substituted, R is H or an alkyl or aralkyl group, X is H or a univalent substituent, Y is H or a univalent substituent, particularly a halogen These dyes are described in Ger 445,270

Printing. I. G FARBEINING A-G (Albrecht Schmidt and Ernst Pfeffer, inventors). Ger 515,305, Jan. 5, 1929 Prints obtained from vat dyes with the aid of methyl cellulose are fixed, without steaming, by means of a salt bath contg alkali and a reducing agent, e. g., hyposulfite Examples are given

Coloring textiles. HENRY DREYFUS. Fr. 694,750, April 29, 1930 The solidity of dyes on textiles, etc., particularly anthraquinone derivs contg NH_2 , alkylamino or arylamino groups, is improved by incorporating in the textiles before, after or during their dyeing, one or more urea bases such as diisomethylthiourea or tetramethylthiourea. A dispersion agent and a protective colloid may also be added Fr. 694,751 describes a similar process in which one or more alkylated diarylamino bases such as tetraalkyldiaminodiphenylmethane (or ethane) are used Cf C A 25, 1685

Colored reserves in textile printing. I G FARBEINING A-G Brit. 336,968, July 22, 1929. Colored reserves are obtained by incorporating with the reserve, com

26—PAINTS, VARNISHES AND RESINS

A. H. SABIN

Desirable properties in paints for toys. W. ADLER *Farbe u. Lack* 1931, 65, 76-7.
—Suitable pigments of a wide range of colors for use in aq. and non-aq. media are discussed. G. G. SWARD

Adhesion in the painting and in the gluing of wood. F. L. BROWNE *Ind. Eng. Chem.* 23, 290-1 (1931). —To improve materially, the durability of house paints on wood means must be discovered for making coatings adhere permanently so that they will wear away instead of falling off in pieces (cf. *C. A.* 24, 2313). Scientific approach to the problem requires a study of the nature of the adhesion between coatings and wood. There are essential similarities between coatings clinging to wood surfaces and layers of glue joining surfaces in glued wood joints. Knowledge already gained about the adhesion between glue and wood (*C. A.* 21, 3425-23, 1304) may therefore serve as a useful guide in the study of adhesion between coatings and wood. F. L. BROWNE

Notes on lead tungstate. H. A. GARDNER *Am. Paint & Varnish Mfrs.' Assoc. Circ.* No. 377, 133-4 (1931). —Pb tungstate as a vehicle for quick-drying paints may be prep'd by adding 30 parts litharge or 35 parts basic Pb carbonate to 100 parts of tung oil fatty acids at 300° F. The paste so prep'd is thinned while hot, preferably with solvent naphtha. Pb tungstate so prep'd. is sol. in a no. of common solvents while hot but gradually pptn from these solns. on standing after cooling. Apparently its pptn does not preclude its use in paints. G. G. SWARD

Protection of surfaces with paint and varnish. ASSEK *Farbe u. Lack* 1931, 78 — An address. G. G. SWARD

Criticism of accelerated weathering tests. PAUL NETTMANN *Farbe u. Lack* 1931, 77. —The usual accelerated weathering test is conducted without proper consideration of the various destructive forces. Until these forces are evaluated, success in accelerated weathering tests will not be attained. G. G. SWARD

Quinquennial review of the mineral production of India for the years 1924 to 1928. BARYTES. A. M. HERON *Records Geol. Survey India* 64, 324-8 (1930). Mineral paints. E. H. PASCOE. *Ibid.* 400-11. —Red and yellow ochre are the only important mineral pigments in India. ALDEN H. EMERY

Bodied pitchard oils. STANLEY A. LEVY. *Am. Paint & Varnish Mfrs.' Assoc. Circ.* No. 377, 139 (1931). —Some phys. constants of pitchard oil thickened to various degrees are given. G. G. SWARD

Glass color standards for varnish. STANLEY A. LEVY. *Am. Paint & Varnish Mfrs.' Assoc. Circ.* No. 377, 140 (1931); cf. *C. A.* 24, 4408. —The $K_2Cr_2O_7$ and $NiSO_4 \cdot 6H_2O$ values for addnl. intermediate glass color standards are given. G. G. SWARD

Lacquer diluents of the petroleum type. JEFFREY R. STEWART *Am. Paint & Varnish Mfrs.' Assoc. Circ.* No. 378, 143-53 (1931). —The kaun butanol no., the dil. ratio, the soly. in Me_2SO , and the aniline point (crit. soln. temp. in aniline) of 27 different lacquer diluents were det'd. In most cases the last named test may be carried out as follows: Warm approx. equal quantities (viz. 3 cc. aniline and 2.5 cc. diluent) in a 6-in. test tube until miscible. Allow the contents of the test tube to cool in a water bath and note the temp. at which the soln. becomes cloudy. Repeat the process twice with 0.3 cc. addns. of diluent each time. The max. temp. at which cloudiness appears is the aniline point. The aniline point is inversely proportional to the other 3 properties. G. G. SWARD

Alcohols for cutting shellac. STANLEY A. LEVY. *Am. Paint & Varnish Mfrs.' Assoc. Circ.* No. 377, 137-8 (1931). —No differences were noted in the working properties of shellac cut with formula No. 1 and with formula No. 5 alc. contg. alcotate and alcohol. G. G. SWARD

Graphical calculation of the composition of ceramic varnishes. ANGELO CASTIGLIONI. *Industria chimica* 5, 1351-5 (1930). —The quantities of the various constituents in a varnish are represented by the sides of polygons having as many axes as there are constituents, and whose length is proportional to the amts. present. Thus when mixts. of varnishes of known compn. are mixed with those of unknown compn., the compn. of the resulting mix being known, the unknown quantities can be scaled off on the diagram by proportion. A. W. CONTIERI

Yellowing phenomenon in coating compositions. I. Oil vehicles. EDWARD F. MALONEY. II. Water vehicles. F. C. ATWOOD. *Am. Paint & Varnish Mfrs.' Assoc. Circ.* No. 370, 461-73; *Am. Paint J.* 14, No. 52D, 25, *Paint, Oil & Chem. Rev.* 90,

No 17, 80 3, *Oil, Paint & Drug Rep* 118, No 10, 54 5(1930) —The theories of yellowing of oil paints are briefly reviewed, 17 references being given. While a slightly elevated temp. accelerates the yellowing of oil vehicles, temps. below 225-50° do not affect water vehicles. Data are given for the hiding power of 15 common pigments in a casein vehicle (compn. not given). Max. hiding power is sq. ft. per gal. of paint was obtained with lithopone which gave 30½ sq. ft./gallon. Some of the other values were Titanox-II 245, 1 ng. whitening, 16½, barytes, 112. The detns. were made on black and white checkered linoleum (cf *C. A.* 24, 2312). G. G. SWARD

Surface protection of the light metals. TH. LEIB AND PA. KOLBE. *Korrosion Metallschutz* 7, 3-11(1931) —This is mostly a discussion of the problems encountered. The use of the Jiroka-Bengough process and of combinations of oil base and nitrocellulose base coatings is recommended. Data are given showing the resistance of such coatings against accelerated tests and outdoor exposure. The use of oil base coatings, applied in 3 thin successive coatings, is recommended. Among the pigments, zinc chromate and iron oxide were found to be the best. Good results were also obtained with zinc white. Expts. were made with the application of oven drying coatings on untempered laural, in such a manner that the drying and tempering took place simultaneously. The advantage is that parts treated in this manner can be stored for a longer period of time without danger of corrosion. If such parts are to be exposed to sea water, addnl. protective coatings must be applied. LEOPOLD JESSAL

"Improved" linoleum brush-out test for hiding power. STANLEY A. LEVY. *Am. Paint & Varnish Mfrs' Assoc., Circ.* No 377, 135-6(1931) —Results of expts. with a brush out board (cf *C. A.* 24, 2312) whose squares conform in brightness to the specifications of the A. S. T. M. are satisfactory. G. G. SWARD

The nature of the resins in jack pine [*Pinus banksiana*]. HAROLD HIBBERT AND JOHN B. PHILLIPS. *Can. J. Research* 4, 1-34(1931) —Green and seasoned jack pines have an av. crude resin content of 4.58% and an ether sol. content of 3.53%. The value of "total acids" present in the resin was the same in each case, but the proportion of "fatty" acids was greater in green wood, while the unsaponifiable matter was considerably less. The seasoned wood contained a lower percentage of fats than the green wood and a correspondingly higher percentage of resin acids. The amt. of unsaponifiable, polymerized terpene substances was also higher. There was a marked decrease in the amts. of phytosterol and "resene" in the seasoned wood. The isolated resin acids obtained by distn. or esterification contained a high percentage of abietic acid formed by transformation of the original acids. The percentage of natural (pimaric) acids was, however, quite high in the resin acids isolated by petr. ether and recrystd. only twice. A higher percentage of cryst. acids was found in the resin acids from green wood than from seasoned wood, presumably because of the change into amorphous products during storage. The fatty constituents were chiefly members of the unsatd. series, present both as free acids, glycerides, or other esters. The seasoned wood contained much less linolic acid in the glycerides than green wood. Oleic acid was present in about the same proportion in both cases. Linolenic was present in only very small quantity. The free fatty acids had practically the same percentage compn. in both green and seasoned wood. The amt. of the total unsatd. fatty acids was higher in the glycerides and free acids from green wood than from seasoned wood. The latter contained a high percentage (30.2%) of oxidized acidic material in the fatty glycerides, indicating that extensive polymerization, or decompn. of some kind, had taken place in the fats present in the green wood during the time of seasoning. The quantities of essential oil obtained from the resins of each kind of wood were very small, amounting to 1.5 to 4.3% of the total crude resins. The products showed no difference in properties, or variation in amt., with time of storage of the wood. Because of the small quantity available for investigation, only a pinene could be identified. The percentages of phytosterol and resene were extremely small, and were appreciably less in the seasoned wood than in the green wood. The amt. of polymerized terpene material found in the unsaponifiable matter was much higher in the seasoned wood. This was probably due to extensive polymerization of the essential oil having taken place during the storage of the wood. J. W. SHIPLEY

Preparation of synthetic resins from alkali lignin. MAX PHILLIPS AND HERMAN D. WELSH. *Ind. Eng. Chem.* 23, 286-7(1931) —The condensation of lignin with various aromatic amines and with furfural yields sol., dark, fusible resins. Possible uses for the products are suggested. W. A. LA LANDE, JR.

Cr oxide [for use as a pigment] (Brit. pat. 336 671) 18. Asphalt or tar compositions

[for paint] (Fr. pat. 694,789) 20. Nitrocellulose from wood-pulp board [for varnishes] (Brit. pat. 330,235) 23. Organic isocolloids [for varnish] (Fr. pat. 694,339) 10.

Paints. N. V. DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ (Fr. 694,640, April 29, 1930) The agglomerating agent for colors containing metals is used in the form of an aqueous dispersion. An example is given of the mixt. of Al powder in white spirit which is poured into an aqueous dispersion of asphalt and the whole is stirred.

Bituminous paint. CRAIGBANK CHEMICAL CO., LTD., and J. H. TAYLOR (Brit. 336,117, Oct. 26, 1929) A paint comprising bitumen or pitch, a volatile solvent such as turpentine, a pigment or mineral filler such as PbO and a vegetable drying oil such as boiled oil is stated to be improved by the addition of 5–15% of tannic acid.

"Weather-resistant" paints. I. G. FARBENIND. A.-G. (Brit. 336,108, May 2, 1929) Paints such as those formed with the usual drying oils, have incorporated with them such a quantity of a basic pigment such as white lead and minium, as is necessary to combine into soaps all free and combined fatty acids present in the binding agent. Other inert constituents such as colcothar, titanium white, lithopone or blanc fixe may also be added. Examples with details are given. Cl. C. A. 25, 1691.

Anti-rust paint. ADOLF SCHNEIDER (Fr. 694,833, May 1, 1930) Fe or its oxides are heated in the presence of O to 750–1100° for a sufficient length of time to produce the gray highly oxidized compound, which is used in a finely divided state with suitable carriers or binders.

Pigments. TITANIUM PIGMENT CO., INC. (Ger. 516,314, Nov. 11, 1926) See Brit. 281,459 (C. A. 22, 3543).

Zinc colors. COMPAGNIE DES MINES, FORGES ET ACIÉRIES DE VITKOVICE and ADOLF ANDRIOL (Fr. 694,290, Apr. 22, 1930) An app. is described for treating pyrites with gases containing H₂S, for the production of H₂S, which is used for the manufacture of Zn colors from Zn lyes.

Lithopone. SACHTLEBEN A.-G. FÜR BERGBAU UND CHEMISCHE INDUSTRIE (Brit. 330,639, Dec. 5, 1928) For rendering lithopone weather resisting it is quenched, after calcination, in water containing water-soluble sulfates such as MgSO₄ or Na₂SO₄, and if the lithopone does not naturally contain some soluble Ba salts a small proportion of a salt such as BaCl₂ or Ba sulfide may be added to it before the calcination, so that the sulfate treatment may form BaSO₄ around the particles of lithopone. Cf. C. A. 25, 222.

Printing inks and pastes. I. G. FARBENIND. A.-G. (Brit. 337,199, Oct. 31, 1929) The process for preparing printing inks and pastes described in Brit. 307,877 (C. A. 24, 252) is modified by substituting, for the resinous products used, unhardened solid products produced by the condensation of "polybasic" organic acids with polyhydric alcohols and one or more "monobasic" carboxylic acids containing at least 7 C atoms, such as the product which may be prepared by heating together glycerol, linseed oil, phthalic anhydride and colophony dissolved in toluene, which may be ground with Lithol Fast Scarlet R. Various other materials may be added such as soluble cellulose derivatives, solvents, plasticizers, driers, etc.

Treating drying oils. GUSTAV RUTH AKT.-GES. and ERICH ASSER (Fr. 694,074, Apr. 17, 1930) Useful products are obtained by heating drying oils, particularly linseed oil or a mixt. of linseed oil and china wood oil, to about 270–280° until polymerization commences. Air is then bubbled through and the heating continued to 310° or more. Cellulose esters or ethers may be added.

Nitrocellulose varnish. KAZUO TAKEMURA (50% to Kazuo Oiwa), U. S. 1,791,879, Feb. 10. A coating composition which is especially suitable for use on ironwork comprises nitrocellulose together with mixed solvents and about 3% of K₄FeC₆N₆ and pulverized bamboo (suitably about 5%).

Nitrocellulose lacquers, etc. W. J. JENKINS and IMPERIAL CHEMICAL INDUSTRIES, LTD. (Brit. 336,615, April 11, 1929) Compositions for the preparation of enamels or lacquers are made by dissolving water-wet nitrocellulose in a solvent which evaporates in air at normal temperatures and has a boiling point not less than about 120° (such as ethylene glycol monoethyl ether or BuOAc) with the addition of a water-miscible or partly miscible volatile solvent such as industrial alcohol or EtOAc in sufficient quantity to keep the water in solution. A solvent of higher boiling point than the main solvent also may be added such as methylcyclohexanone, and water-wet or moist pigments may be added also, as may ingredients such as dammar gum, tricresyl phosphate and the like.

Coloring nitrocellulose lacquers and plastics. I. G. FARBENIND. A.-G. (Ger. 516,315, Aug. 29, 1929) Addition to 515,057 (C. A. 25, 1692) The dyes obtainable from hydrochlorides of dialkylated nitrosoamines and monoalkylated *m*-aminophenols are used, instead of the dyes specified in Ger. 515,057.

Lacs and varnishes. JOHANNES SCHEIDER (Ger. 513,540, Feb. 21, 1928, Linoleic

acid isomers (octadecadien 9,11-acid 1) obtained from ricinoleic acid are esterified with multivalent alcs., especially glycerol, and the product is used for the manuf. of lacs and varnishes. Resin may be added. Thus, technical ricinoleic acid is heated to 200° and distilled *in vacuo*. The product going over at 260-270° is heated to 200° with glycerol to give a bright oil. Other examples are given. Cf. C. A. 25, 1400.

Coating composition. VICTOR H. TUCKINGTON (to Bakelite Corp.). Can. 308,848, Feb. 24, 1931. Resinous compns are prep'd by reaction of 100 parts of cresol, 200 parts of tung oil and 25 parts of $(CH_3)_3N$ at a temp. of 190-200°. Cf. C. A. 25, 834.

Coating compositions containing m-styrene and softeners. WALTER E. LAWSON (to E. I. du Pont de Nemours & Co.). U.S. 1,792,102, Feb. 10. Compns suitable for coating metals, glass, wood, etc., comprise m-styrene and triarsyl phosphate, dibutyl phthalate, diethyl ethane, glyceryl dibenzyl ether, butyl stearate or homologs of these compds. Cf. C. A. 24, 980.

Coating metals with synthetic resins. HEROLD A.-G. Brit. 337,192, April 30, 1929. Solns. of phenol aldehyde artificial resins are applied to metal surfaces while the latter are at a temp. above the b. p. of the solvent (which may be "spurt") to evaporate the solvent and the coatings are hardened by heating. The rate of hardening may be restricted by the use of incomplete condensation products or by dilg. the coating soln. with cresols, glycerol, glycol or high b. p. polyglycols, or addn. of solids such as feldspar, zircons or kieselguhr.

Purifying and decolorizing rosins. ROBERT C. PALMER and CLYDE O. HENKE (to Newport Co.). U.S. 1,791,633, Feb. 10. Rosin dissolved in a volatile hydrocarbon material such as petroleum naphtha is treated with resorcinol at a temp. (suitably about 90-110°) at which the resorcinol is in liquid form and below the b. p. of the hydrocarbon material, and after agitation, the materials are allowed to stand and the liquid resorcinol layer containing the impurities is removed. Cf. C. A. 25, 424.

Rosin decolorization and purification. DUBAIN C. BUTTS (to Hercules Powder Co.). U.S. 1,791,638, Feb. 10. Rosin is heated in an inert atm. such as CO_2 to a temp. substantially within the range of 230° to 350° while avoiding substantial decompn. or distn.

Inlaid inoleum. W. O. L'ESTRANGE. Brit. 336,277, July 8, 1929. Mech. features.

Condensation products from vinyl esters. I. G. FARBENFABRIK A.-G. Brit. 236,237, Nov. 18, 1928. In the production of polymerization products such as those of vinyl acetate, chloroacetate, chloride and benzoate and of lacquers from the resinous derivs. thus obtained, there is added (before, during or after the polymerization) a small proportion of a basic substance sol. in org. solvents such as guanidine or its derivs., hydrazine or its derivs., ethylamine, formaldehyde or phenylhydrazones and N compds. having plasticizing properties (such as a condensation product of urea and formaldehyde) also may be added. Benzoyl peroxide also is added in polymerizing vinyl benzoate. The production of lacquers is described. Cf. C. A. 25, 339.

Molded articles from synthetic resin compositions. BAKELITE CORP. Brit. 336,265, July 10, 1929. Articles with a raised surface design are made by applying a colored mixt. (such as a plastic phenol resin and bronze or other pigment) to a phenol resin compn. in a partially cured state and then completing the curing. Various details of procedure are described.

Colored molding mixture containing a synthetic resin. BAKELITE CORP. Brit. 336,264, July 10, 1929. A mixt. comprising a potentially reactive resin, a filler such as wood flour and a pigment is subjected to extrusion whereby the colored particles are distributed along definite flow lines.

Molding articles of artificial resin. HEROLD A.-G. Brit. 337,184, Oct. 24, 1928. Articles of artificial resin such as balls, tubes or dishes are cast by the use of a cast mold of Pb or other easily fusible metal within an outer mold which may be made by either casting or dipping. Various structural details are described.

Modifying resins, waxes and fatty oils. HERBERT HÖNELL. Austrian 120,679, May 15, 1930. Resins, waxes and fatty oils are heated with a subordinate amt. of a primary condensation product obtained under alk. conditions from CH_2O and a phenol having a hydrocarbon substituent in the p-position to the OH group. In this way, the resins and waxes are rendered harder and of higher m. p., and fatty oils are made more viscous. Thus, an oily condensation product from p-amyphenol and CH_2O may be introduced into 5 times its wt. of a fused coumarone resin m. 50-55°, the temp. being gradually raised to 200°. A clear resin m. 70-75° is obtained. Other examples and addnl. details are given.

Synthetic resins. IMPERIAL CHEMICAL INDUSTRIES, LTD., and R. HILL. Brit.

336,698, Aug 3, 1929 Blending of monohydroxylated fatty oils such as drying oils with resins such as those formed from glycerol and phthalic anhydride or similar components is facilitated by replacing some or all of the polyhydric alc. by a reaction product of a polyhydric alc. with a lower aliphatic aldehyde such as CH_3O , e. g., glycerol and polymerized CH_2O are condensed together, then phthalic anhydride and tung oil or linseed oil, or phthalic anhydride, tung oil, colophony and glycerol are added and the mixt. is heated until a homogeneous resin is produced. Varnishes prepd from such products dry rapidly in the air or when heated and their drying may be facilitated by adding substances such as Co linoleate.

Synthetic resins. BAKELITE CORP. Fr. 694,047, April 16, 1930. White masses of compns. having a basis of phenol resinoid are made by the reaction of phenol about 100 with CH_2O about 200 and di-Et phthalate 70 parts in the presence of oxalic acid acting as catalyst. Water is removed from the reaction mass, a volatile solvent is added and the whole mixed with "albalith". The liquid mixt. thus obtained is applied to a filling material. The app. used must not contain Fe or Cu. Cf. C. A. 25, 1400.

Synthetic resins. EDGARD ISRAËL. Fr. 695,036, Aug 7, 1929. Urea and CH_2O are condensed with glucose or lactose as condensing agent, in the presence of a metallic oxide, preferably ZnO , and polymerized at 80° .

Artificial resins of the glycerol phthalic anhydride type. BAKELITE CORP. Brit. 336,645, July 10, 1929. A "polybasic" acid such as phthalic anhydride or succinic acid and a polyhydric alc. such as glycerol are heated together until "slag formation" begins and the reaction is then arrested by rapidly cooling the mass throughout so that transformation to a "slag condition" is prevented (as by casting in thin layers). The product may subsequently be baked for hardening or may be broken up, remelted and cast. Dyes may be added and various details of temp., employed, etc., are given.

Artificial resins. COMPAGNIE FRANÇAISE POUR L'EXPLOITATION DES PROCÉDÉS THOMSON-HOUSTON. Fr. 694,669, Apr 29, 1930. Artificial resins are prepd by the reaction of amines such as PhNH_2 with insufficient aldehyde, then adding more aldehyde and heating under pressure to cause the resin to pass to the infusible state.

Artificial resins. IMPERIAL CHEMICAL INDUSTRIES, LTD. Fr. 694,181, April 18, 1930. Rapidly hardening resins are made by the union of a polybasic acid polyhydric alc. resin with a condensation product of CH_2O and urea or thiourea or with substances capable of giving a product of this kind, this union being effected in an alc. solvent and with the aid of heat. As solvents ethylene chlorohydrin, BuOH or hexahydrophenol may be used. Cf. C. A. 24, 740.

Artificial resins. LOUIS C. F. PÉCHEN. Fr. 694,978, May 5, 1930. Artificial phenol aldehyde resins are hardened by heating them under vacuum to $95\text{--}105^\circ$ in vessels placed in an oil of low d., and then to $140\text{--}180^\circ$ under atm. pressure still in the oil. Cf. C. A. 25, 1692.

27—FATS, FATTY OILS, WAXES AND SOAPS

E. SCHERUBEL

Determination of the fat content of oil seeds. M. MONHAUPT. *Chem.-Ztg.* 55, 70(1931).—M. modifies Dopfer's method for the detn. of fat in oil seeds as follows: place 4 g. of the ground seed in a round bottom, heavy glass bottle together with 40 cc. of gasoline h. $40\text{--}60^\circ$, 10–12 g. of sand and 5 or 6 steel balls of 10–12 mm. diam. and grind in a shaking machine for 1–2 hrs. or until apparently "dust-free", allow to settle, evap. 20 cc. of the gasoline ext., weigh and calc. the % of fat according to the formula: $x = (100 \times 40 \times f) / (4(20 - 1.0753 f))$, in which f is the wt. of the oil residue from 20 cc. gasoline soln., and 1.0753 is the sp. gr. factor for oil (1.00/0.93).

The determination of unsaturation of fats and acids. III. Wijs iodine number. J. VAN LOON. *Chem. Umschau Fette, Oele, Wachse u. Harze* 37, 257–62(1930), cf. C. A. 24, 3663.—The Wijs method for detn. of I no. shows the theoretically true unsatn. value when correct exptl. conditions are selected, except with triple bonds which become only $\frac{1}{2}$ satd. by halogens. There is no substitution by I even in prolonged action of the reagent. For official analyses the I no. of the acids (after removal of the unsapon matter) should be detd. in addn. to that of the oil, because their I no. is more const. than that of the oil. If the total acids show an I no. which increases on continued action of the reagent, a polymerized product or some unknown acid may be present. The Wijs method satisfies all scientific requirements and is applicable to general fat chemistry.

P. ESCHER

Some physical constants of fatty acids. (Specific heat, viscosity, absorption of

ultra-violet light.) E. L. LEOXER. *Allgem. Öl- u. Fettig* 27, 237-41 (1930) —The sp heats of stearic, lauric and oleic acids were measured at several temps and calcd for others. A new formula for calcn of viscosity, based on thermodynamic principles, is derived $\log \eta = (q^*/4.571 T) - 2.75 \log T + ET + C$, where η = coeff of viscosity (in poises), T = temp, q^* = heat of assocn at 0° (cal per mol), E = const, C = const. The values of q^* , E and C for the following substances are, resp., butyric acid 1960, 0.001770, 2.9932, valeric acid 2330, 0.001663, 2.9070, capric acid 2740, 0.001401, 2.8200, heptylic acid 3210, 0.001470, 2.5791, caprylic acid 3475, 0.001576, 2.4883, nonylic acid 3655, 0.001696, 2.4160, capric acid 3700, 0.001902, 2.3885, myristic acid 5030, 0.001670, 2.0221, stearic acid 5490, 0.001907, 1.7020, pentane 724, 0.002097, 3.0077, water 5825, 0.008222, —1.9613. The values of η for butyric acid at 8 temps between 0° and 160° were calcd and found to agree within 2.2% with the values given in the tables of Landolt Börnstein. Measurements of the absorption of ultra violet light were made on the satd fatty acids in the liquid state, rather than in soln. In making the measurements a photoelec. Na cell was used instead of a photographic plate. In this way it was possible to make more accurate measurements than heretofore. It was found that all the fatty acids observed showed sp absorptions of the wave lengths between 270 and 280 $m\mu$. This sp absorption was observed in synthetically prepd, as well as in naturally occurring, fatty acids but not in triglycerides. It was further observed that the sp absorption for the shorter chain acids was for the shorter wave lengths. Thus the sp absorption range was between 270 $m\mu$ for propionic acid, and 280 $m\mu$ for stearic acid. The cause for this sp absorption has not yet been definitely established. W. F. BOLLENS

Method for estimation of total unsaturated fatty acids of higher unsaturation than oleic acid in fats and oils. N. DASKIN. *Allgem. Öl- u. Fettig* 27, 243-4 (1930) —The method is based on the conversion of unsatd fatty acids of high mol wt into ether acids (cf. C. A. 24, 2142). The saponid fat carrying excess alkali of 0.1% is polymerized in an atm of H_2 for 2-3 hrs at a temp of 290°. This procedure reduces the unsatn of those fatty acids contg more than 1 double bond to that of oleic acid. The difference in I nos of the sepd fatty acids before and after polymerization is called the "Mutit" no and is a measure of the fatty acids contg more than 1 double bond. The "Mutit" no of several oils is given: olein (tech.) 12, peanut oil 23, cottonseed oil 50, soy bean oil 50, sunflower seed oil 40, linseed oil 60, herring oil 54, cod liver oil 74.1. W. F. BOLLENS

Characteristic fatty acids of vegetable and animal fats with reference to the biological families in which they occur. T. P. HILBREN. *Allgem. Öl- u. Fettig* 27, 219-22, 255-8 (1930) cf. C. A. 24, 4176 —A compilation is made of the percentage fatty acid compn of a large no of fats and oils arranged according to the biological families in which they occur. W. F. BOLLENS

Manipulation and evaluation of the Mackey test. K. KERN. *Seifensieder-Ztg* 58, 29-32 (1931) —K summarizes former exptl results as follows. The real cause of the temp increase in the Mackey test is the rapid oxidation of the linoleic or linolenic acid when present in the olein to be tested. This temp increase is accelerated by iron soaps or other pos catalysts and is inhibited by neg catalysts like β naphthol. The presence of soaps should be confirmed by an ash detn (upper limit less than 0.07%) and by a qual test for Fe. When present, the temp test should be repeated after removing the Fe soap from the sample by extrn with dil HCl. The I no and Kaulmann's thiocyanate no should be detd to detect any presence of linoleic acid if a neg catalyst is suspected. The presence of much unsaponifiable matter also reduces or retards the temp increase. The standard limit of 100° or 150° after 1 hr should be modified, since oleins which contain acids with more than 1 double bond show a rapid temp increase after $\frac{1}{4}$ hr, the cotton being more or less carbonized. Directions for making the test are given in much detail. P. ESCHER

Further notes on the Mackey test. H. WOLF AND R. HEILINGÖTTER. *Chem. Umschau Fette, Öle, Wachse u. Harze* 38, 24-5 (1931), cf. C. A. 25, 429 —The selection of the material for the cotton bearing cylinder may well be left to the (German) Olein Commission. The samples of olein should be tested in the same condition as received, it would be illogical to attribute the temp rise due to the presence of any soap to the presence of "unsatd fatty acids" in an otherwise unobjectionable sample. The soaps cause a temp rise only after prolonged (1.5 hrs) heating at 100°, while unsatd acids usually show a temp rise at a low temp, after $\frac{1}{4}$ hr. An olein becomes a fire hazard through the presence of unsatd acids, not by the presence of soaps. The present tendency is to reduce the time of heating to 1 hr rather than to increase it to 3 hrs. The moment the thermometer in the cotton registers 70° should be considered the start-

ing time of heating. It is not necessary to pass a stream of air through the app., since the natural ventilation through the tubes is sufficient. P. TSCHIR

The pyrolysis of vegetable oils of high acetyl value. RAYMOND DELANY AND RAYMOND CIBARRONAT. *Bull. sci. pharm.* 38, 17-23 (1931). A. E. MEYER

Yield differences in the oil mill. E. I. BITTER. *Chem. Umschau Fette, Ole, Wachse u. Harze* 38, 33-4 (1931). —B calls attention to unaccounted for differences (about 1%) in the oil mill between the calculated yield based on lab. analysis and the actual yield in the factory. It is suggested that this difference is due to phosphatides or lipoids showing different solubilities under lab. and factory conditions. P. TSCHIR

Analysis of refined (olive) oils. GASTANO MARCONA. *Atti III congresso naz. chim. pura applicata* 1930, 864-8. —From chem. and phys. tests of a great many samples of olive oil it is concluded that refined oil from the first extr. acts like a true olive oil as regards chem. and phys. const. and various analytical tests. The exceptions are the Wood light test and the Belter test. Oils from the second extr. show a reduction in av. quality and chem. and phys. tests. The characteristics of edible olive oil should be standardized strictly. Oils which have a n_D^{20} above 1.4628, a thermal index above 47, unsaponifiable content above 1.5%, and which do not correspond to the Belter, Blarez and Pichini-Morawsky tests should be excluded. Addn. of refined first-extr. olive oil should be limited, so as not to vary these tests. I. M. SWARTZ

Correlation of oil content to presence of color in the stem of castor (*Ricinus communis*). G. B. PATWARDHAN. *Poona Agr. Coll. Mag.* 22, 16-8 (1930). —In 32 samples of beans from red stemmed castor plants the oil content ranged from 40.1% to 54.97% with an av. of 47.91%, while in 24 samples from green stemmed plants the oil content ranged from 35.75 to 51.21% with an av. of 51.22%. The analyses were made on seed of parents which had proved true to stem color in the succeeding generations. K. D. JACON

Detection of castor oil and peanut oil in abrasin oil. HANS WOLFF AND J. RANINOWICZ. *Farben Ztg.* 36, 590-7 (1930). —Castor oil can be detected by the formation of octyl alc. on fusion with KOH. The amt. can be calcd. from the acetyl values of the sample and of an H_2O (80-85%) ext. Peanut oil can be detected by the formation of arachidic acid and idetd. by means of its n_D^{20} , which is 1.469-72, that of abrasin oil is 1.517. G. G. SWARN

Suggestions for the proposed unified methods of analysis of Turkey-red oil and similar products. WIZDER. *Chem. Umschau Fette, Ole, Wachse u. Harze* 38, 31-7 (1931). —This is the text for the proposed standard methods of analysis as worked out by the German Commission for Fat Analysis jointly with the Intern. Soc. of Leather Chemists, the National Comm. of Trade Specifications and the Union of Turkish Red Oil Mfrs. The methods are based on extensive analytical investigations and on control checks by public analysts, but cover only sulfonated oils whose SO_3 can be completely split off by boiling HCl . P. TSCHIR

Constituents of the leaves of pine (*Pinus Thumbergil*). I. JENJIRO SAKURAI. *J. Pharm. Soc. Japan* 51, 31-9 (1931). —S. isolated from the petr. ether ext. of the leaves of pine a wax-like substance (I) which had the following phys. properties: d_4^{20} 0.9201, m 78-80°, acid no. 29.81, sapon no. 218.90, ester no. 189.05, sapon no. after acetylation 219.18, and I no. 7.55. Sapon. of I (100 g.) gave mixed solid fatty acids (II) (85 g.) II was esterified and fractionated into the following constituents: lauric, palmitic, stearic, hydroxypalmitic acids. F. I. NAKAMURA

Textile soap. G. T. HALBERSTADT. *Proc. Am. Assoc. Textile Chem. Colorists* 1931, 67-70. *Am. Dyestuff Repr.* 20, 109-12 (1931). —A review. M. H.

Lime-proof soaps. W. SCHRAUTH. *Seifen- u. Fett-Ztg.* 58, 61-3 (1931). —The usual lime-proof soaps are the Na salts of sulfonated fatty acids which keep the Ca soaps, formed by hard water, in such a finely divided state that they are not deposited upon the fabric of the washed goods. S. describes a new type of soap like salts made by reducing the higher fatty acids to their alcs., sulfonating them and converting the sulfonates into their Na salts of the general formula $\text{RCH}_2\text{SO}_3\text{Na}$. These salts have the properties of soaps, excelling in emulsifying power, lauric Na sulfate shows with Traube's surface tension pipet 29 drops in spindle oil against 17 drops for regular coconut-oil soap. They are not easily hydrolyzed and are very stable toward Ca and Mg salts. P. TSCHIR

Rancidity of soaps and its relation to the properties of fats. F. WITTEK. *Seifen- u. Fett-Ztg.* 58, 3-5 (1931), cf. *C. A.* 25, 613. —W. prepd. soaps from various fats and examd. their keeping quality toward oxidation in the presence and absence of catalysts and light. Of the hardened fish oils whale oil resists oxidation most, herring oil is equiv. to edible tallow, sardine oil turns brown on exposure to light or to catalysts, the latter

2 oils on standing take on an intense odor of hardened oil, peanut, sesame and palm-oil are sensitive to air and light. Coconut oil and its soap keep longest, soaps from palm-kernel oil and Babassu fat do not keep quite as well. Soaps from corn, grape kernel and linseed oil soon turn rancid even without catalysts. Soaps with 15% or less of WW or FG resin show little change, but larger quantities make them unstable and sticky.

P. ESCHER

Moisture determination in soaps and fats by distillation. E. SCHLECKER. *Seifen-Industrie* 21g 58, 66-7 (1931) —From lab. expts. S. concludes that H_2O detns. in soaps by the distn. method with petr. benzene yield low results, unless a sufficient amt. of olein (20 g. per 30 g. soap flakes) has been added before distn.; addn. of dry sand is superfluous. Butter needs no olein addn., and correct results are obtained by direct distn. with petr. benzene. S. prefers xylene in place of petr. benzene, since the latter holds some of the distd. H_2O in min. suspension.

P. ESCHER

Indirect estimation of potassium and sodium in filled potassium-sodium soaps. G. ENIGGE. *Algem. Öl- u. Fettg.* 27, 223-5 (1930) —The indirect method of Davidsohn (*C. A.* 24, 3553) is proved with a soap of known compn. Ground pumice is used as the filler. The results check very well with known compn. and also with results obtained by the perchlorate method.

W. P. BOLLEYS

Surface tension of alkaline soap solution. SUTZU OGURI. *J. Soc. Chem. Ind. Japan* 33, Suppl. binding 471-4 (1930) —A modified Millard's app. was used to det. surface tension by the drop-wt. method. All measurements were made at 40°, and the tip of the pipet had a radius of 3.23 mm. The surface tension was calcd. from Tate's law $2\pi r\gamma = mg$, in which $2\pi r$ = the circumference of the tip, γ = surface tension, m = effective mass of the drop, and g the acceleration due to gravity. One table and 3 graphs are given.

E. SCHREIBER

The effect of cleansing agents on linen and cotton (SINCLAIR) 25. Determination of vegetable oils in butter (DE SUTTA VETRA) 12. Structure of synthetic mixed triglycerides (BRATTACHARYA, HILBITH) 10. Apparatus for extracting gases from oils (Fr. pat. 653,565) 1. Fractional condensation of oil vapors (Ger. pat. 513,513) 22. Residue of sulfonic acids from oil-refining [fat-splitting agents] (Fr. pat. 654,235) 18. Modifying waxes and fatty oils (Austrian pat. 120,579) 26. Continuous production of dispersions of waxes (U. S. pat. 1,792,557) 22.

Extracting oils and fats. C. SCHLOTTERHOFF and H. BRANDT (trading as Schlottterhose & Co.) Brit. 237,579, March 29, 1929 The extn. process described in Brit. 216,881 (*C. A.* 24, 2562) is mod'ed by subjecting the material to be extd. to a treatment with the vapor of the solvent, or direct or counter-current, in addn. to or instead of treatment with liquid solvent. Various details of app. and procedure are described.

Extraction of oil, etc., from soy beans. M. SATO and C. ITO. Brit. 335,273, June 10, 1929 Extn. with a solvent such as EtOH is effected at a temp. above the b. p. of the solvent but below 120° (suitably at about 85°), and the soln. of fatty oil is drawn off and cooled to below 50° to sep. oil and obtain a liquid residue contg. coloring matter, sugar and lecithin. Brit. 335,274 (M. Sato and Y. Ishida) describes the extn. of soy beans and the like with a solvent such as alc. admixed with benzine or benzene of such concn. and quantity that the water in the resulting liquor is less than 10%. Extn. is effected above 65° and oil is sep'd by cooling. An arrangement of app. is described.

Extracting oleaginous materials with solvents such as in degreasing bones. JOSEPH SAVAGE. U. S. 1,791,555, Feb. 10 A volatile hydrocarbon solvent is circulated in a closed circuit through an extn. chamber, passing through the chamber as vapor, under superatm. pressure, the vapors are condensed, extd. water is automatically sep'd., and hot sep'd. solvent is returned to the circuit at a temp. near the b. p. of the solvent at the pressure existing in the circuit. App. is described.

Glycerol distillation. I. G. FARRINGTON A-G. Brit. 235,508, July 12, 1929 Crude glycerol is treated, at its b. p. under the reduced pressure employed (suitably 20 mm. or 50 mm.) with finely divided liquids (such as wet steam or a mist of CCl_4 or benzene assocd. with air or CO_2) inert to glycerol and having a b. p. not exceeding that of water and assocd. with gaseous or vaporous carriers (the operation being continuous). Various details and examples are given.

Stabilizing fatty substances and soaps. R. T. VANDERKILT CO., INC. Fr. 654,286, Apr. 22, 1900 Fatty substances and soaps are stabilized by the addn. of a phenolate prep'd. by the reaction of a phenylphenol such as 2- or 4-hydroxyphenyl with a base which is itself a stabilizer such as diphenylguanidine, triethanolamine or an alkylene-

diaryldiamine. If the fatty substance or soap is animal or vegetable a *p*-phenylphenol is used, if mineral an *o*-phenylphenol is used. Cf. C. A. 25, 430

Soaps. A. IMHAUSEN. Brit. 335,954, May 6, 1929. Soaps of high lathering power comprise grain soap and coconut oil or palm kernel oil paste soap contg. little salting-out agent. Various details of manufacture are given.

Soap. ALBERT A. BESNARD. Fr. 695,028, Aug. 6, 1929. A soap which cleans the hands without water contains agar agar 2, psidium 3, glycerol 50, Na_2CO_3 50, black soap in pulp 50, " NH_3 " 25, Javel water 5 and distd. water 815 parts.

Soap. CARLOS CONEN. Fr. 694,220, April 19, 1930. A soap is made by adding naphtha benzine or refined petroleum to a fatty acid and adding an alk. substance such as Na_2CO_3 in the dry state.

28—SUGAR, STARCH AND GUMS

J. K. DALR

Defecation of cane sugar juices. NATHAN LEVY W. *Bull. assoc. chim. suc. dist.* 47, 502-3 (1930), cf. C. A. 4, 3310—The author reviews his method for detg. the right amt. of lime for defecation of cane juices. F. CAMPS-CAMPINS

The use of lime and sulfurous acid in the defecation of sugar juices. NATHAN LEVY W. *Bull. assoc. chim. suc. dist.* 47, 500-2 (1930)—The harmful effects of too much or too little lime in the defecation of cane sugar juices are reviewed. If excess lime is added and then followed by SO_2 to neutrality, all the evil effects of excess lime are produced and the subsequent SO_2 does not destroy them. The author advocates sulfitation before, instead of after, liming. A slight excess of SO_2 is not harmful; the right amt. of lime, as detd. on raw juice, is then added. There will be no deleterious result even though the resultant juice is slightly acid. F. CAMPS-CAMPINS

Advances in beet-sugar manufacture in 1930. EDMUND O. VON LIPPMANN. *Chem.-Ztg.* 55, No. 17, *Fortschrittsber.* No. 1, 33-6 (1931). E. H.

The "Werkspoor" rapid crystallizer. R. BONATH. *Intern. Sugar J.* 32, 306-9 (1930)—The earlier method of operating crystallizers was in series, utilizing the surrounding air for cooling. Further developments were made with cooling jackets, and later fixed internal tube systems were adopted. A more successful type was developed by Lafeuille, who introduced a fixed tubular system, around which the hermetically closed cylinder contg. the massecuite rotated. The requirements of a crystallizer are (1) shortening the period of fermentation, (2) avoidance of formation of false grain, (3) simple and substantial design. The "Werkspoor" crystallizer consists of a trough which may be either U- or bottle-shaped. The revolving water-cooled stirring system consists of a hollow drum, of large section, riveted to which are a number of semi-circular hollow cooling pockets coupled by suitable pipe connections inside the central drum. The transport of the massecuite from one end to the other of the crystallizer is effected by a screw-shaped element placed outside and independent of the cooling pockets. W. L. OWEN

Starches and the constitution of starch. ST. V. NÁRAY-SZABÓ. *Z. physik. Chem.*, Abt. A, 151, 420-4 (1930), cf. C. A. 25, 16—A discussion of recent work. W. L. HILL

The effect of plowing-under cane trash upon the available N of the soil (OWEN, DENSON) 15. Potash fertilizers at the "Belle Etoile" distillery (DUBOSC) 15. An inexpensive crystal or molasses separator (KING) 1. Sugar anhydrides (MICHAEL, MICHAEL) 10. N-containing sugars (BERTHO, *et al.*) 10. Continuous production of dispersions of gums (U. S. pat. 1,792,067) 22. Extracting [sugar from beet roots] (Fr. pat. 694,602) 17.

Device for cleaning sugar canes. RALPH S. FALKNER. Ger. 513,928, Mar. 2, 1929.

Apparatus for storing sugar beet in water without loss of sugar. FIRMA R. FÖLSCH. Ger. 513,927, Dec. 10, 1929.

Apparatus for mixing liquids with solids, particularly for liming sugar juices. GILCHRIST CO. Ger. 516,359, July 19, 1925. Brit. 247,541 (C. A. 21, 668).

Refining sugar. RAFFINERIE TRILEMONTAISE, SOC. ANON. Ger. 516,360, Oct. 1, 1926. See Brit. 278,302 (C. A. 22, 2680).

Carbohydrate derivatives. I. G. FARBENIND A.-G. Fr. 694,696, April 30, 1930. Carbohydrate derivs. which are sol. in water and org. solvents and which are useful

for the prepn of films, lacquers, artificial silk, plastic substances, swelling and impregnating agents, etc., are made by treating a carboxylate or one of its derivs with an alkylene oxide or one of its homologs or analogs in the presence of an aq alkali. Examples are given of the treatment of potato starch, dextrin or cotton with ethylene oxide, propylene oxide, etc.

Dextrose. CORN PRODUCTS REFINING CO. Ger 513,570, July 22, 1923. Anhyd. dextrose is prepd from converted starch solns by crystn at 57-49°. The crystals are sep'd from the mother sap by centrifuge. Cf C A 24, 4652.

Soluble starch. ROBERT HAAKE and WALTER HAAKE. U. S. 1,702,088, Feb 10. A highly conc'd oxidizing liquid such as NaOCl soln is added in small quantities to the starch to be disintegrated and mixed with it while maintaining the starch in its powdery condition, thus elevating the heat of reaction to such an extent that the chem. transformation is performed without any residue and a sufficient evapn is effected of the slight surplus of moisture contained in the treating soln added.

Starch products. INTERNATIONAL PATENTS DEVELOPMENT CO. Ger. 516,261, July 3, 1928. See Brit 204,979 (C A 23, 2065).

29-LEATHER AND GLUE

ALLEN ROGERS

Tannery effluent. L. Effect of various gases on the nitrogen distribution. EDWIN R. THIES and PHILIP KRATZ. *Ind Eng Chem* 23, 69-71 (1931).—The effect of p_n value on evolution of gases from soak waters contg. NaNO_3 was that gases were evolved at p_n values 1 to 5, and absorbed at higher p_n values. Bubbling O_2 through soak water promotes formation of N_2 , indicating deamination, it causes the effluent to become very colloidal in character and retards the formation of volatile sulfides. H_2 furthers the formation of H_2S , causes deamination with subsequent reduction, and finally causes the formation of volatile acids. H B MERRILL.

Fat-liquoring of chrome leather. Effect of hydrogen-ion concentrations on oil adsorption. EDWIN R. THIES and FRANK S. HUNT. *Ind Eng Chem* 23, 50-3 (1931).—The amt. of oil taken up by chrome leather is studied for numerous mixts. of sulfonated and row oils and com. fat liquors, as a function of the p_n value of both the skin and the fat liquor. Characteristic curves are obtained for each oil mixt. employed. Relation of the strength of the leather to the kind of fat liquor and to the amt. of oil adsorbed is shown. The adsorption of oil is shown as "oilation" by means of the dilatometer. H B MERRILL.

A note on the cause of certain red colorations on salted hides and a comparison of the growth and survival of halophilic or salt-loving organisms and some ordinary organisms of dirt and putrefaction on media of varying salt concentrations. MADGE E. ROBERTSON. *J Hyg* 31, 84-93 (1931).—Certain brick-red stains on salted hides are produced by the growth of halophilic organisms which come from marine salt used in curing the hides. There are many varieties of these organisms which flourish best in high concns. of salt and fail to grow below 6%. The organisms of dirt and putrefaction though certain of them survived as high a concn. as 30% for a considerable time, failed to multiply above 8%. The non halophilic cocci survived at high salt concns. much longer than the non halophilic bacilli. JOHN T. MYERS.

Glue making from chrome-leather scrap by the chromic hydroxide process. C. STRIEPEL. *Chem Ztg* 54, 804-5 (1930).—In this process the chrome complex in the leather is changed to a stable, non tanning hydroxide by "intensive liming." The process is similar to the magnesia process, but the yields obtained with the latter are inferior. The process is far superior to the various acid processes, is economical as regards time, labor and equipment, and yields a glue of good quality. H B M.

Protecting leather from fungi (Brit. pat. 335,244) 18

Treating hides with liquids for tanning, etc. CHARLES WESLEY NANCE. Ger 513,929, Dec. 31, 1927. See Brit. 310,823 (C. A. 24, 749).

Treatment of skins before tanning. DAVID L. LEVY. Fr 604,407, April 24, 1930. Skins are treated with a soln. of mineral or org. acids and alkali, alk. earth or metallic salts before the usual tanning operation. Cf C A 24, 1247.

Tanning agents. I. G. FARRENHEND A-G. Fr 604,257, April 22, 1930. Condensation products which are sol. in water and have very pronounced tanning and chag.

ing properties are prep'd by the reaction of urea, its homologs or substitution products and aliphatic aldehydes on di- or poly hydric phenols of the C_{11} series, their homologs or substitution products, preferably in the presence of acid condensing agents. Part of the phenols may be replaced by tanning substances of vegetable origin. Several examples are given. Cf. C. 4 25, 1116.

Tanning compositions. W. SAILER. Brit. 336,984, April 18, 1929. Sulfate cellulose lvs. phenols, cresols and natural tanning substances such as a bark ext. are caused to react with enzymes such as phenolases, peroxidases or dihydrazes and with substances yielding O such as H_2O_2 or perborates. The products may be further treated with halogens and are suitable for use preliminary to chrome tanning.

Greasing leather. HERMANN BOLLMANN and BRUNO REWALD. Ger. 516,187, Sept. 7, 1927, 516,188, Oct. 23, 1927, and 516,189, Dec. 23, 1927. Addns to 514,899 (C. 1 1704). There is used (1) a mixt of lecithin and train oil emulsified in a relatively large amt of water (516,187), or (2) an aq. soln of lecithin (516,188) or (3) a mixt or emulsion of lecithin and fatty oil with a small quantity of soap or sulfonated oil, with or without a small quantity of mineral oil (516,189).

Dyeing leather. I. G. FARBENING A. G. (Bartholomäus Vossen, inventor). Ger. 513,842, Feb. 4, 1923. Leather is dyed by acid azo dyes built up by the process amine \rightarrow 1-amino-8-hydroxynaphthalene-3,6-disulfonic acid \rightarrow resorcinol \rightarrow β -nitraniline, the amine indicating sulfanilic, picramic or naphthionic acids. The example describes the dyeing of deacidified chrome leather by the dye built up from picramic acid.

Machine for mordanting hides. VEREINIGTE HUTSTOFFWERKE BLOCH & HIRSCH, C. F. DONNER G. M. B. H. Ger. 513,698, Mar. 23, 1929.

Coated fabric (imitation leather). ARTHUR W. BIRWELL (to Alox Chemical Corp.). U. S. 1,791,711, Feb. 10. A coating compn. suitable for use in the manuf. of artificial leather" comprises nitrocellulose and a softening agent consisting essentially of the water insol., sat'd aliphatic, high mol. wt., carboxylic acid product obtainable by contacting a free-O-contg. gas such as air with a petroleum hydrocarbon mixt consisting of a 45-43° B ϕ petroleum distillate in the liquid state, in the presence of an "oxidation extender" such as Aln oleate at a reactive temp. not above about 155° (suitably about 140°) and at a superatm. pressure but not greater than about 350 lb. per sq. in., to the point of incipient formation of petroleum insol. compds. in the mixt. An app. is described suitable for oxidation of the petroleum distillate.

30—RUBBER AND ALLIED SUBSTANCES

C. C. DAVIS

The oil and the lipase-like enzyme in Para rubber seed. YOSHITORA IWAMOTO. *J. Soc. Chem. Ind., Japan* 33, Suppl. binding 409-11 (1930).—Fallu seeds from Malaya were used in the expts. The following data give the kind of seed, the grade of oil from each kind, the color of the oil, its taste and its acid no., resp.: fresh with yellow white kernel, 1st, faint yellow, sweet, 7.21; older with dirty yellow kernel, 2nd, yellow, bitter, 24.71; partly putrefied, 3rd, brownish yellow, bitter, 77.10; badly putrefied with dark gray kernel, 4th, brown, very bitter, 110.86. The results show that fresh seeds are suitable for practical oil extn. Oil obtained by cold pressing is superior to oil by hot pressing, as shown by a higher yield, a lighter color, a sweeter taste and a much lower acid no. (19.70 vs. 45.22). A bitter taste is best removed by NaOH. Oil extd. with petr. ether from fresh seeds showed d_{4}^{15} 0.9234, n_D^{20} 1.4757, acid no. 7.12, I no. 138.8, solidifying point 2°, sapon no. 189.20, Reichert-Messel no. 2.28, acetyl no. 2.41, unsaponifiable matter 1.62, insol. hexabromide (fatty acids) 15.4%. By the Pb salt Et_2O method, the fatty acids yielded 79.99% of yellow liquid acids and 17.81% of white solid acids. The former when oxidized with alk. $KMnO_4$ by the Hazura method gave dihydroxystearic, sativic and humic acids. This shows that the liquid acids contain considerable oleic, linoleic and linolenic acids. After recrystn. from 95% $EtOH$ at -20° , the solid acids yielded a white cryst. powder which m. 62° and had a neutralization no. 204.62, which was probably a mixt. of 70% stearic acid and 30% palmitic acid. The high acid no. of the oil from decayed seeds suggested the presence of lipase. The following data give the acid nos. of the oils extd. with Et_2O from 4 samples before and after 92 days at room temp., resp.: (1) seeds with husks, 5.41, 15.72; (2) crushed kernel with husk removed, 6.30, 48.62; (3) crushed kernel with oil removed and then ground with rapeseed oil, 1.20, 93.37; (4) same as (3) but

the spaces between the threads of the lining, to reduce the stiffness of the boot and to prevent adhesion to the last. Various details of manuf. are described.

Use of forms in making balloons, gloves, coats or other dipped rubber goods from aqueous dispersions of rubber. ISAAC W. ROBERTSON (to Thermo Process Co.) U. S. 1,792,187, Feb. 10. Various details of procedure are described.

Puncture-sealing mixture for tires. P. G. MASTERS. Brit. 335,980, July 5, 1929. A mixt. for placement in inner tire tubes is formed of castor oil 0.5 pint, talc 1.5 oz., cellulose 1.5 oz., such as wood dust, water 0.5 pint, gum acacia 1 oz. or less, with or without a small addn. of benzene, "petrol" or rubber soln.

Rubber insulating material. WM. HAYAN WIEGAND. Fr. 694,333, April 23, 1930. Insulating materials contg. rubber have added thereto C black retaining its adsorbent power in amt. up to 10% of the wt. of rubber or 3% of the total mass. An example of such a material contains rubber 984.6, ZnO 30, C black 29, monosulfide of tetramethylthiouran 3.5 and S 40 parts. Other examples are given.

Mercaptothiazole derivatives. WM. P. TEN HOUWER (to Rubber Service Laboratories Co.) U. S. 1,792,090, Feb. 10. In producing products such as various rubber vulcanization accelerators, reaction is effected between an amine such as piperidine or diphenylguanidine and the product obtained by the reaction of a salt of mercaptobenzothiazole with a Cl deriv. of CS₂. Cf. C. A. 24, 2470.

Elastic material. OCTAVE PIRARD. Fr. 694,998, Aug. 2, 1929. A light elastic material is made, e. g., of wood charcoal 35, bark fiber 10 and pure rubber 55%.

Transparent sheets comprising synthetic rubber. I. G. FARBERNINO A-G. Brit. 336,075, Sept. 20, 1929. Reinforced sheets are made by applying a soft coating of a polymerization product of butadiene or its homologs to an open meshed fabric (such as tinned iron wire mesh or a vegetable fiber material) and hardening the coating (as by the action of O or by heating). Condensation products such as those derived from oxylene or other aromatic hydrocarbons with butadiene in the presence of a condensing agent such as AlCl₃ also may be used.

Synthetic rubber. I. G. FARBERNINO A-G. Brit. 336,339, Aug. 2, 1929. Removal of polymerization products of diolefins from the reaction vessel is facilitated by lining the vessel with readily removable materials such as paper, cardboard, textile fabrics, leather, or films of sheet metal or cellulose derivs. or polymers of diolefins or of rubber (which latter may be vulcanized or otherwise suitably treated), or waxes, readily fusible alloys or bitumen. Various examples and details of app. and procedure are given.

Synthetic rubber. I. G. FARBERNINO A-G. Brit. 337,019, May 25, 1929. In polymerizing diolefins such as butadiene by the action of alkali or alk. earth metals or their mixts. or alloys, in the presence of org. solvents such as ether, the solns. of the polymerization products are continuously removed from the catalyst as soon as the polymerization is effected, as by a flowing stream of solvent or soln. The products may be unsuitable for the production of soft rubber, but are suitable for the manuf. of various hard materials.

Vulcanizing synthetic rubber. I. G. FARBERNINO A-G. Brit. 335,970, July 2, 1929. Polymerization products such as those of butadiene or its homologs or analogs are vulcanized in the presence of an accelerator previously dissolved in a solvent or admixed with a liquefying agent, which gives a product of improved stretch and breaking strength. Solvents of various kinds may be used such as water, alc., hydrocarbons, chlorinated hydrocarbons, ethers, weak acids and liquid bases. Stearic acid may be used as a softening agent. Cf. C. A. 24, 3400.

Rubber vulcanization accelerator. WINFIELD SCOTT (to Rubber Service Laboratories Co.) U. S. 1,792,041, Feb. 10. The reaction product of formaldehyde with crotonaldehyde dianilide is used as an accelerator. Cf. C. A. 25, 618.

Rubber vulcanization. ALBERT A. SOMERVILLE (to R. T. Vanderbilt Co.) U. S. 1,791,876, Feb. 10. In effecting vulcanization with the aid of an active accelerator such as mercaptobenzothiazole a small quantity of a peroxide such as BaO₂, PbO₂, CaO₂, or SrO₂ is added to inhibit or prevent scorching or prevulcanization. Various examples are given.

Apparatus for vulcanizing rubber coatings on hollow metal articles such as press-rollers of paper-making apparatus. CHRISTIAN H. GRAY. U. S. 1,792,298, Feb. 10. Structural features.

Forming vulcanized joints in rubber articles such as air-tubes, rubber balls or other hollow products. DUNLOP RUBBER CO., LTD., and G. R. MEAD. Brit. 336,076, Sept. 21, 1929. Mech. features.

CHEMICAL ABSTRACTS

Vol. 25.

MAY 10, 1931

No. 9

1—APPARATUS AND PLANT EQUIPMENT

W. L. RADGER

Filtering disks of sintered Pyrex glass. W. F. BRUCE AND H. E. BENT. *J. Am. Chem. Soc.* 53, 990-2(1931).—A layer of powd. glass (60-, 80- or 100-mesh), 1.5-2 mm. thick, is sintered in a Ni tube of suitable diam. About 2 min. at a bright red heat is required. The disk obtained may be sealed to a Pyrex tube. B. A. SOUTER

Further applications of the centrifugal filtration tube. EYALD L. SEAR AND LOUIS F. ROWE. *Ind. Eng. Chem., Anal. Ed.* 3, 147-8(1931), cf. *C. A.* 23, 5355.—This device can be used to det. the approx. amt. of solid and liquid in a mixt. at any definite temp. Hence binary f. p. diagrams may be constructed or a eutectic compn. detd. B. A. SOUTER

Portable electric stirrers in the chemical industry. HANS WOLLENBERG. *Chem.-Ztg.* 54, 804(1930).—A small, compact, well protected elec. stirrer which can be clamped on any vessel is described. A. L. HENNE

A modified slow-combustion pipet. KENNETH A. NOBB. *Ind. Eng. Chem., Anal. Ed.* 3, 159 (1931).—A modified Stalene pipet for slow-combustion analysis is described, and the advantages of this method over the explosion method are pointed out. D. F. BROWN

Wash bottle for organic, poisonous or corroding fluids. WALTER FERMATIN. *Chem.-Ztg.* 54, 804(1930).—To the conventional wash bottle a vent tube is added. The vent is closed with a finger, while the liquid is being forced out of the bottle. To interrupt the liquid stream the finger is lifted before blowing is stopped. This prevents disagreeable fumes from backing into the operator's mouth. A. L. HENNE

Comparative efficiencies of gas-washing bottles. F. H. RHODES AND D. R. RAGSTRAW. *Ind. Eng. Chem., Anal. Ed.* 3, 143-4(1931).—Each bottle was filled to the normal working height with NaOH (32.7 g./l.). Then air contg. 13.4% CO₂ was passed through at increasing rates of flow until that rate was found at which some CO₂ escaped absorption. The order of decreasing efficiency is: Friedrichs spiral, Schott & Gen. No. 101, No. 88c, No. 83, Muencke, Habermann. The order is approx. the same when 3.1% CO₂ is used. B. A. SOUTER

Direct-reading volume-meter. M. T. CARLISLE AND RHEBA USTER. *J. Chem. Educ.* 7, 152(1930).—This vol.-meter for irregular solids consists of a cylinder 4.8 cm. in diam. and 12 cm. long, the bottom of which is drawn out into a tube 10 cm. long contg. a bulb of 10 cc. capacity. A glass stopper with a finely graduated outlet tube, 30 cm. long and 5 mm. in diam., is inserted in the top of the cylinder. A leveling bulb of Hg is attached to the bottom tube. The cylinder and bottom tube are filled with water, after which the leveling bulb is brought up to a mark on the bottom tube. The level of the column of water forced up into the upper outlet is then read and the leveling bulb lowered again. After the test specimen is dropped into the cylinder, the operation is repeated and the difference in water levels noted. ANN NICHOLSON HIRD

Potentiometer for the determination of μ . N. N. KAVTSEVICH. *Dokl. Akad. Nauk SSSR*, 1, 152(1930).—Potentiometric measurement of the e. m. f. is discussed, and a simplified potentiometer which gives readings with an accuracy of 0.01 μ is described. The novel feature of the potentiometer is the introduction of a telephone in place of the galvanometer. J. S. JORRE

Colloid mills. AUGUST CHWALA. *Kolloid-Z.* 54, 243-51(1931); cf. *C. A.* 25, 1122.—A review of German patents. ARTHUR FLEISCHER

An automatic constant-level device for liquid air. R. B. SCOTT AND F. G. BRICK. *Rev. Sci. Instruments* 2, 171-2(1931). E. H.

Apparatus for the continuous extraction of large quantities of material at high temperatures. FRANK WRENDS. *Biochem. Z.* 231, 173-4(1931).—An app. capable of extg. 5 kg. or more of material has an outer vessel that contains a heating liquid (liquids of various b. ps. can be used for extra. at diff. temps.) and an inner container for the ma-

terial and the solvent. The solvent is boiled in a flask outside the container, so connected that the solvent circulates through the material and then returns to the flask.

S MOROULIS

Extraction apparatus. R C BAKER. *J Soc Chem Ind* 50, 54T(1931)—A drawing showing a modified Soxhlet extractor is given, and the advantages possessed by the app over the usual type are presented.

N A LANGE

Note on centrifugal dust extractors. M SEILLAN. *Chaleur et ind* 10, 289-90 (1929), cf *C A* 24, 992—S discusses the trajectory of an element of gas emerging from the rotor of a centrifugal pump and draws the parallelogram of velocities at the exit of a ventilator. This is followed by a discussion of cyclone extractors.

S L B ETHERTON

An apparatus for determination of true vapor pressure. E J SQUIRE. *Petroleum World* 28, No. 1, 66-70, 101-9(1931)—An app for detg true vapor pressure should be designed so that the sample may be introduced as it exists at the point of sampling, so that vapor pressures may be detd at different temps, and so as to be speedy enough in operation for routine testing. The app here described can be said to meet these conditions and is suitable for liquids that do not attack Fe, Hg or glass, are not too viscous to be properly agitated, do not have an abs vapor pressure over 2000 mm of Hg, and do not contain emulsion.

R W KELLY

Apparatus for purifying gases. Z. VERESS. *Magyar Chem Folyoirat* 35, 155-7 (1929)—The current of gas drives the solvent into an absorption tower, whence it returns dropwise into the original container.

B C A

The technic of gas analysis. A J ANTHONY. *Z. Biol* 90, 633-6(1930)—An app equipped with shaking attachment is described. The specific difficulties for nitrous oxide analysis are discussed.

FRANCES KRASNOW

Automatic analyzers of gas mixtures based upon thermal conductivity of gases and their industrial use. DE FOSBARTVILLE. *Chaleur et ind* 11, 469-71(1930)—A discussion of an article by Jarrier (cf *C A* 25, 1413), in which F meets the weaknesses of app described by J. PAUL JARRIER. *Ibid* 471-2—J replies and gives more detail.

S L B ETHERTON

A new titration colorimeter. K. MAYER. *Biochem Z.* 231, 314-6(1931).

S MOROULIS

Helical pumps and their use in industry. ERNST BLAU. *Chem.-Ztg* 54, 801-2 (1930)—Descriptive.

A. L. HENNE

Improvements in the construction and use of the Berthelot-Mahler calorimeter. RICCARDO DE BENEDETTI. *Rivista tec ferrovie ital* 38, 58-60(1930)—Several details of the mech parts of the app are modified, and methods of calibrating and operating are suggested for serial work and for the detns with liquid combustibles.

G A B

Generator for the production of carbon dioxide of high purity. EDGAR J POTH. *Ind Eng Chem, Anal Ed* 3, 202-3(1931)—The construction of a generator for CO₂ or other gases is described, for which the following advantages are claimed: (1) It is compact and is constructed entirely of glass without ground-glass joints. (2) It can be warmed and pumped free of occluded gases. (3) It has a novel automatic feeding arrangement. (4) Reagents are completely used, and none of the gas produced is wasted.

CHANNING WILSON

Flow divider for fractionation under diminished pressure. A M SCHWARTZ AND M T BUSH. *Ind Eng Chem, Anal Ed* 3, 138-9(1931)—In vacuum distn, total condensation and return of part of the condensate as reflux permit complete and easy control of the reflux ratio. A stillhead with a Hg check valve has been found suitable for this purpose. Constructional details are given.

B A. SOULE

Glass temperature and float regulators. D F OTHERM. *Ind Eng Chem, Anal Ed* 3, 139-43(1931)—Constructional and operating details are given for a controller of pressure or temp when a flowing liquid causes the change of conditions. The sensitive element is a vapor pressure bulb that operates a "flow diverter." A temp, never over 0.2" high, may be maintained for weeks. The device is recommended for lab or semi works operations. Several float valves are also described.

B A. S

Glass electrode and vacuum-tube potentiometer. DONALD H CAMERON. *J. Am Leather Chem Assoc* 26, 7-23(1931)—Advantages and limitations of the electrode are discussed. In tannery analyses, the chief advantage is freedom from poisoning and disturbance by oxidation-reduction reactions, the chief disadvantage, aside from the rather laborious standardizations required, is the failure to function at pH values higher than 9. Different forms of electrode are described. MacInnes' form, using the Ag/AgCl electrode, is preferred. A vacuum-tube potentiometer circuit, on the Wheatstone bridge principle, is described in detail.

H. B. MERRILL

A vacuum-tube method of temperature control. FRANCIS O. SCHMITT AND OTTO H A SCHMITT *Science* 73, 289-90(1931)—The app is essentially a type FG-27 thyatron regulator tube with a toluene-Hg regulator inserted in the grid circuit. A large un-insulated metal water bath has been held for days to within 0.005°, well insulated baths may be held within 0.001°.

J H. MOORE

Something new in vacuum drying chambers. HOLLAND-MERTEN *Chem.-Ztg.* 55, 153-4(1931)—A rectangular chamber, built by Vacuumtrockner, Erfurt, is described. It is made up of independent heating elements, any no of which may be used, the sides of the elements forming the walls of the chamber. Internal dead space is eliminated, internal condensation and steam leakage are minimized, and less wt. and floor space are required for a given capacity.

J. H. MOORE

The determination of the heat of wetting. WALTER REIMER *Wiss Arch Landw., Abt A, Pflanze* 4, 383-91(1930)—A new calorimeter for detg the heat of wetting is described. The heat tone is measured by the expansion of toluene, which surrounds the inner wall of the calorimeter and connects with a graduated capillary tube. A sensitivity of 0.25 cal is possible. Details of the construction and use of the calorimeter are given.

W GORDON ROSE

Chemical engineering memoranda. XIII. Further notes on heaters, including air heating by steam. CHARLES H. BUTCHER *Ind Chemist* 7, 95-6(1931); cf. C A, 25, 364.

E. H.

A very sensitive, self-registering heat-flow meter. GEORG HOFBAUER *Arch Wärmewirt* 10, 398-402(1929)—The instrument, which is intended primarily for measuring heat flows through the walls of buildings, consists of 5 multiple couples, each contg 100 Fe-constantan elements, with junctions on opposite sides of a rubber sheet 0.3 cm thick, which is protected by thin outer rubber sheets and surrounded by a protecting ring. Good accuracy can be attained with flows as low as 1.5 cal per sq cm per hr.

ERNEST W. THIELE

The energy problem of electric heating. F. MERKEL *Arch Wärmewirt* 10, 407-8(1929)—A general discussion of heat pumps with special reference to their use in distn.

ERNEST W. THIELE

Packed heat from a new type of electric steam generator. HOWARD BIDWELL AND LINUS H JONES *Ind Eng Chem* 23, 298-300(1931)—The problem of obtaining large quantities of high pressure steam for short periods for sterilization work has been solved by the construction of an elec. steam generator of original design. The boiler of the generator is offset at an angle, almost the entire vol of water in the generator and reservoir being made available for steam. Immersion heating elements are used, the size and energy consumption depending on the pressure and quantity of steam required. The design incorporates the safety features of small diams. Sterilizing pressure may be obtained within 30 sec from the time of closing the autoclave.

A S S

A practical hot stage for the microscope. EDWARD H WALLACE AND MARY L. WILLARD *J Chem Education* 8, 706-11(1931).

E H

Modern electrically heated aluminum-fusing furnaces. KIRCHRATH *Metallbörse* 20, 2731-2, 2779-80(1930)—Furnaces, especially those built by Gautsch and Brandt, with Ni-Cr heating elements, are described. The elements are either bands or spirals placed in recesses in the roof so that heat is radiated down on the charge.

J. H. MOORE

Determination of the local strength of fabrics, felts, paper, rubber, film, wood, leather, metal, etc. II. FRIEDRICH SCHUBERT *Chem.-Ztg* 55, 134-5(1931); cf. C A, 25, 1708—The app referred to in C A 23, 1187, 4103, has demonstrated its value. It is described, its use and the results are discussed and a sample calcn is given.

J H. MOORE

Early steps in the development of the Columbia soil rod. BUELL B. LEGG *Gas Age-Record* 67, 111-8(1931)—Work is being done on the development of a soil rod which will enable the determination of the galvanic action of the soil, the specific resistance, the tendency to polarization and the potential difference between adjacent soils. Preliminary work indicates a close relationship between results obtained with such a rod and corrosion of test samples. The rods at present have a copper electrode and a steel electrode suitably insulated. The bottom, steel electrode is tapered to provide a driving tip.

LESLIE B BRAGG

Modification of Ostwald's electric thermoregulator. E BUTTERWORTH AND D A DERRETT-SMITH *J Sci. Instruments* 7, 233(1930)—Change in calibration due to allowing the instrument to cool is avoided by omitting the 2 side tubes of the standard form, sufficiently fine adjustment being obtainable by means of the screw in the terminal

head It is advisable to distil the toluene over Hg to insure complete removal of S

B. C. A.

A laboratory thermostat for continuous operation. P. W. SCHENK. *Ber.* 64B, 368-70(1931) —A thermostat is described which will operate without special attention for a long period of time with an accuracy of 0.02° A Dewar tube is surrounded by a wooden box which contains the stirring motor, battery, relay, condenser and Hg switch. A particular feature of the thermostat assembly is the inclosed Hg switch operated by a relay and making and breaking the current to two cylindrical heating bulbs in the thermostat. The relay is actuated by a Hg-in-glass thermoregulator. A condenser of 1-2 mf is used across the regulator. For temps below room temp a const. stream of water is passed through a Cu U-tube in the thermostat, and the same heating system is used to maintain the lower temp.

ALLEN S. SMITH

Simple universal thermostat for laboratory fermentations, refractometry and pycnometry. H. FINK. *Wochschr. Bran* 47, 461-4(1930) —This app consists of a eudiometer with capillaries and containers for fermentation, a thermostat with heating elements, thermoregulator and stirrer, and an electrically driven shaking device and fillings to accommodate 20 pycnometers and 12 capsules mounted for reading by an immersion refractometer. Six or 12 fermentation flasks can be used at once. These are provided with gas measuring fittings. The bath remains const within 0.2° and is used within the range 15° to 40°.

S. JÖRGA

Oven-temperature control and design of a bi-metal thermostat. W. E. BENTON. *Gas World* 94, 35-7(1931) —A math treatment of the fundamental principles of bi-metal thermostat design.

F. H. BERGMAN

A cooling unit for low-temperature thermostats. T. J. B. STIER. *Science* 73, 288-9(1931) —A Cu funnel is fed with cracked ice from an insulated hopper, with means for drawing off water as the ice melts. The funnel is immersed in the thermostat to the proper level, found by trial, and will hold the app const to ±0.01° at room temps. It costs about \$1.00. Details are given.

J. H. MOORE

Ag and its application to chemical plant (McDONALD) 9. Colorimeters, spectrophotometers and nephelometers (DICAUD) 7.

Colorimeter JOHN C. BAKER and CHARLES F. WALLACE (to Wallace & Tiernan Products, Inc.) U. S. 1,794,134, Feb 24.

Immersion pyrometer for fused metals. RUDOLF HASE. *Ger.* 514,682, May 11, 1929. Details of construction are described.

Optical absorption-pyrometer KAISER-WILHELM INSTITUT FÜR EISENFORSCHUNG E. V. (Gerhard Naeser, inventor). *Ger.* 517,023, July 9, 1929.

Thermometer MAX F. MOELLER (to Moeller Instrument Co.). U. S. 1,793,930, Feb 24.

Thermometer connected with an indicator at a distance. HERMAN SCHLAICH. U. S. 1,794,324, Feb 24. Structural features.

Distance signalling thermometer suitable for use in stacks. D. G. ZEITLIN. *Brit.* 338,267, Sept 14, 1929. Structural features.

Thermometer or like device containing a gallium alloy. SYLVESTER BOVER (to General Elec Co.) U. S. 1,793,303, Feb 17. Ga or alloy which may contain Ga, Sn and In is placed in a tube of material such as fused SiO₂ to which the alloy does not adhere.

Viscometer HARRY W. KLINGER (to Hercules Powder Co.) U. S. 1,793,807, Feb 24. X rays are utilized for observing the passage of a ball or the like through opaque liquids. *Cl. C. A.* 25, 3.

Air filter LOUIS KLAFF (to American Air Filter Co.). U. S. 1,794,115, Feb 24.

Air filter. ALBERT BEYSEL-HÖSSY. *Swiss* 142,423, Oct. 4, 1929. A labyrinthine passage for the air causes deposition of foreign bodies.

Air filter for liquid-storage vessels. KARL PRILLINGER and FRANTISEK PRILLINGER. *Swiss* 142,109, Sept. 11, 1929.

Fabric air filter. C. G. VOKES. *Brit.* 337,966, Nov 1, 1929. Structural features.

Oil filter. CHARLES W. MCKINLEY (to A. C. Spark Plug Co.) U. S. 1,792,854, Feb 17.

Unactivated "prefiltering medium" suitable for treating oils, sugar solutions, etc.

JOHN J. NALGLE. U. S. 1,794,862, March 3. A material of low decolorizing and deodorizing value, but which has a highly porous structure permitting a high rate of filtration, comprises carbonized lignin residues substantially free from all alk. residues.

Filter cloths. VICTOR ROEDER Ger 517,260, Jan 18, 1929 Peripheral reinforcing means for perforations in the cloth is described

Apparatus for filtering liquids under pressure. J HOFFMANN and T KRIBERNIGG Brit 337,818, Aug 14, 1929 Structural features

Rotary, impeller agitating device and filter for treating liquids such as sugar solutions or oils. JOHN J NAUGLE U S 1,794,916, March 3 Structural features

Magnetic separator. MITSUO KOTZUKI Ger 517,259, Apr 8, 1925 Corresponds to Brit 254,030

Means for separating pneumatically conveyed materials from the conveying medium. WILHELM KNOLLE Ger 517,050, Jan 17, 1928

Separator for oil and water, etc. IGNAZ MÖLLER Ger 517,425, Apr 20, 1929 See Austrian 115,650 (C A 24, 2286) and 117,500 (C A 24, 3403)

Apparatus for separating oil from compressed air. MARTIN J BINKLEY (to Calor Oil Burner Corp) U S 1,793,460, Feb 24 Structural features

Settling apparatus suitable for the separation of solids from liquids and gases. WALTON C GRAHAM, HORACE S RUMSEY and ASHUR U WETTERBERG (to Gilchrist & Co) U S 1,794,374, March 3 Structural features

Spray-system washer for separating dust from gases. DAVID S JACOBS (to Babcock & Wilcox Co) U S 1,793,620, Feb 24 Structural features

Röntgen-ray tubes. W D COOLIDGE (to British Thomson-Houston Co, Ltd) Brit 338,430, Jan 16, 1929 Structural details of tubes which may have a Mo plate of sufficient thickness to assume a uniform temp and a surrounding focusing member of Ni or Mo

Quartz rod (with a cone-shaped depression in its end) for use with ultra-violet lamps. CLEON W SYMONDS. U S 1,794,557, March 3

Thermionic cathode. SANDOR JUST (to General Elec Co) U S 1,794,298, Feb 24 A cathode is formed of a metal such as Mo having a m p above 1000° and an electron emissive material comprising a mixt of oxides including an alk earth oxide and a rare earth metal oxide, in which the latter is present to only about 0.1% the quantity of the former

Electron-discharge apparatus. DUDLEY A MULLANEY (to General Elec. Co). U S 1,794,315, Feb 24 A control electrode for a space current device has its surface mechanically roughened to decrease emission of secondary electrons from it

Electron-discharge devices. ÉTABLISSEMENTS INDUSTRIELS DE L C GRAMMONT ET DE A. GRAMMONT Brit 337,971, Dec 10, 1928 Electron emitting filaments (which may be formed of pure or oxidized W or of W coated with oxidized metal such as Fe, Ni or Cu and may if desired be wound with fine taping to retain the coating) are coated with alk earth metal oxide by depositing on the filament a single or double cyanide of the alk earth metal and decomposing the cyanide to form the oxide after the filament is mounted in the electron discharge device. Various details of procedure are described

Electron-discharge devices. H. C. RENTSCHLER (to Westinghouse Lamp Co), Brit. 337,704, Feb 16, 1929 Oxidation of the surface of the anode may be effected by cooling in air after baking in H, and selective reduction of the oxides of Ni and Fe may then be effected by H at 800–1000° A small proportion of H₂O vapor in the H prevents the formation of nitride or carbide of Cr due to the presence of any N or said hydrocarbons in the gas used For further details see U S 1,760,525 (C A 24, 3405).

Discharge device for compressed gases such as carbon dioxide. RENEE M. LEMOINE U S 1,794,185, Feb. 24. Plugs of porous material such as tow impregnated with a liquid of low f. p. such as alc. are placed in expansion devices at the points which are most exposed to freezing U. S. 1,794,186 relates to an app for atomizing paints, insecticides, etc., with CO₂

Gas-analysis apparatus. J G DALOZ Brit 338,410, June 20, 1929 Various structural details are described of an app suitable for detn of fire damp The app may be attached to a miner's lamp

Automatic gas-analysis apparatus. C A HARTUNG Brit 337,862, Aug 21, 1929 Various details are described of app in which the gas to be tested, such as flue gas, is treated with another gas such as H₂ and reaction effected as for the purpose of detg residual O in flue gases.

Apparatus for determining and registering the density of gases and liquids. SIGMENS & HALSKE A-G. (Josef Kröaert and Heinz Grüss, inventors) Ger. 517,364, Mar. 11, 1928. The detn depends on the velocity at which the gas or liquid flows through a narrow orifice.

Apparatus for carrying out catalytic reactions between gases under high tempera-

ture and pressure. MARIA CASALE-SACCHI Swiss 142,144, Apr 5, 1929. Details are given

Apparatus for the catalysis of gas. SOC ANON D'OUGRÉE MARDIAYE Swiss 141,303, Mar 4, 1929 The app is used for such catalytic reactions as the oxidation of NH_3 into oxides of N

Apparatus for maintaining a desired relation between the pressures of gas and air, etc. A C IONIDES Brit 337,935, Oct 28, 1929 Various structural details are described

Device for regulating the quantity and composition of gas mixtures such as those used for anesthesia. ANTON GRAMBERG (to I G Farbenind A-G) U.S. 1,793,608, Feb 24 A device is described for operating needle valves controlling the supply of the gases

Apparatus for treating gases with liquid sprays for washing, saturating, cooling or other purposes J A REAVELL Brit 337,507, Aug 22, 1929 Structural features

Gas washer, with rotary perforated washing drum. ZSCHOCKER-WERKE KAISERSLAUTERN A-G (Ernst Heilmann, inventor) Ger 517,352, Oct 26, 1929

Device for mixing liquids with added substances, for chemical reactions. EUGEN STEURER Ger 514,937, June 9, 1929

Centrifugal atomizer for liquids. AKTIESELSKABET NIRO Ger 514,714, July 18, 1929 The app comprises a rapidly rotating plate and is used for atomizing liquids preparatory to their reaction with gases

Centrifugal drum apparatus for clarifying liquids. H LIST Brit 337,628, Nov. 22, 1929 Structural features

Apparatus for clarifying aqueous liquids, etc., by gravity separation. FRITZ MIENER, U.S. 1,792,919, Feb 17 Structural features

Apparatus for extractions with liquid solvents LE ROY C. TRESCOTT (to Contact Filtration Co) U.S. 1,794,874, March 3 Various details of construction are described of an app for extrn on the Soxhlet principle, such as extrn of clay used in filtering oil An agitating device is mounted in the extrn chamber

Valve for regulating the flow of liquids or liquefiable solids such as sulfur, wax, fat, pitch, etc. ARTHUR T PRENTICE (to Imperial Chemical Industries, Ltd) U.S. 1,793,747, Feb 24 Structural features

Noncorrodible container for liquids WILHELM BESCHETANICK, Ger 514,538, Nov 27, 1929 The container is built up of sections coated with Pb, Cu, Al or other non-corrodible plates

Evaporator with steam-heated tubes suitable for concentrating liquids. FRANZ HOFMEISTER and HANS MELHARDT (to Chemische Fabriken J Bellak) U.S. 1,793,174, Feb 17 Structural features

Evaporating apparatus KARL FRIED Ger 517,170, Jan 8, 1928 The app. combines evapn in thin layers with multiple-effect vacuum evapn

Hollow-plate heating apparatus for evaporating liquids, etc. SOC ANON POUR L'IND CHIM A BALE, Brit. 338,097, April 29, 1929 Structural features

Multistage steam drier. SCHMIDT SCHNEIDER & CO (Otto H Hartmann, inventor) Ger 514,883, Dec 17, 1929

Centrifugal drier with a filter drum. I G FARBENIND A-G and FRITZ SANDER Ger 514,568, July 23, 1927 Addn to 401,162

Apparatus for drying steam. CHRISTIAN HOLSMEYER Ger 517,402, May 10, 1928

Apparatus for drying sheet materials. COW MFG CO Brit 337,432, May 29, 1929 Structural features

Apparatus for drying materials by circulating air which is alternately heated and cooled. JOHAN G OLSSON and FRANS I E STENFORS U.S. 1,795,094, March 3. Structural features

Apparatus and procedure for spray desiccation of materials such as blood, milk or soap. FRED F PEASE (to F F Pease, Inc) U.S. 1,794,978, March 3 Structural and mech details.

Sealing ring for rotating cylinder apparatus such as kilns and driers. BURNETT E GREEN (to Huron Industries, Inc.) U.S. 1,794,454, March 3

Rotary cooling drum (for cooling material discharged from rotary furnaces or kilns). H STEINMANN (to F Krupp Grusonwerk A-G) Brit. 338,069, Jan 18, 1929

Atmospheric water spray cooling tower. CARL F. BRAUN (to C F Braun & Co) U.S. 1,794,394, March 2

Vapor-liquid contact column suitable for distillation, dephlegmation, absorption,

etc. CLAUDE B. SCHNEIBLE (to trustees of Joseph Schneible Trust). U. S. 1,794,986, March 3

Column-and-superposed-tray apparatus for removing carbon by washing from bubble towers used for oil distillation. GUY B. BOGART (to Texas Co.). U. S. 1,792,887, Feb. 17 Structural features

Surface condenser suitable for use with steam. PAUL A. BANCEL (to Ingersoll-Rand Co.) U. S. 1,794,135, Feb. 24 Structural features

Apparatus for aerating water with carbon dioxide. H. ZWICKY Swiss 141,040, Sept. 24, 1929

Apparatus for proportionate mixing of chemicals with water or other fluids supplied through pipes. JAMES W. PARKER, SABIN CROCKER and JAMES H. WALKER. U. S. 1,792,757, Feb. 17 Structural features

Condensing apparatus suitable for use with steam engines or turbines. RAYMOND P. MOORE (to Ingersoll-Rand Co.) U. S. 1,793,119, Feb. 17 Structural features

Tiltable pan for receiving ammonium salts from saturators. FIRMA CARL STILL. Ger. 517,476, July 7, 1928

Lip-tilting crucible furnace. J. H. LUDLOW and ASSOCIATED ELECTRICAL INDUSTRIES, LTD. Brit. 338,013, Nov. 23, 1929

Muffle furnace. ZAHN & CO. G. M. B. H. Ger. 514,502, May 31, 1927 A muffle furnace with a heat resisting hearth for the production of sulfate from NaCl and H₂SO₄ or bisulfate, is described

Horizontal or muffle furnace suitable for treating sand to remove its iron content. V. A. COLLINS. Brit. 337,845, Aug. 13, 1929 Structural features

Annealing furnace. VEREINIGTE DRAHTWERKE A.-G. (TRÉFILERIES RÉUNIES S. A.). Swiss 138,657, May 10, 1929 Details of preheating and cooling are given

Chamber furnace with regenerative chamber for preheating air and gas. C. OTTO & CO. G. M. B. H. Ger. 507,912, Feb. 28, 1929

Two chamber furnace. C. OTTO & CO. G. M. B. H. Ger. 517,175, Mar. 5, 1927. Addn. to 511,515 (C. A. 25, 1127)

Regulator for inclined-grate furnace. FIRMA M. STREICHER. Ger. 514,663, July 6, 1928

Traveling-grate furnace with compressed air admittance below the grate. L. & C. STERNICLER. Ger. 514,843, July 19, 1929

Oil-fired furnace. ERNST FELDHOFF (Erich W. Becker, inventor). Ger. 514,521, Dec. 7, 1928

Gas-fired boiler with plate-like members. STREBELWERK G. M. B. H. Ger. 514,873, Apr. 19, 1929

Gas-fired angular chamber kilns. LUDWIG RIEDHAMMER. Ger. 517,334, Mar. 1, 1929 A method is described of attaining uniform temp. in the upper and lower parts of the combustion chambers

Gas burner. CHARLES J. SCHIFFERLE (to Harry R. Struthers). U. S. 1,794,868-9, March 3

Gas burner. J. W. BAKHUIZEN. Brit. 338,345, Nov. 5, 1928

Safety pilot-jet system for gas burners. A. C. BECKER. Brit. 338,122, April 13, 1929 Two pilot jets and a main burner and associated devices are so arranged that in the event of one of the pilot jets being extinguished it is reignited from the other. Brit. 338,124 and Brit. 338,128 also relate to similar devices

Furnace wall with parallel tubes for cooling. OSCAR NAGAARD U. S. 1,793,004, Feb. 17

Means for heating or intensively cooling hollow rolls or drums for working rubber, etc. PAUL TROESTER, MASCHINENFABRIK. Ger. 517,361, July 9, 1930

Heat-exchange apparatus suitable for use as a condenser. GEORGE T. JACOBS (to Alco Products, Inc.). U. S. 1,794,336, Feb. 24 Structural features

Heat-exchange apparatus suitable for heating air for combustion by hot gases. JEAN B. E. E. ARNOULT (to Établissements Ernest Arnould). U. S. 1,794,573, March 3 Structural features

Apparatus for electric heating of liquids in receptacles by immersed electric resistance devices. W. WINSBUR. Brit. 337,762, May 3, 1929 Structural features

Heating and cooking vessels. "SPINX" SPOJENÉ SMALTOVNÁ TOVÁRNY NA KOVŮVÉ ZBOJÍ AKC. SPOL. Brit. 337,313, Dec. 24, 1929 An inner vessel such as one of rustless steel is fitted within an outer sheath or jacket of metal of higher heat cond. such as Al or Cu Various structural details are described

Apparatus suitable for heating chemicals in large kettles, etc. SIDNEY E. MEYER (to National Aniline & Chemical Co.) U. S. 1,792,918, Feb. 17 Various details are

described of a heating system by which a large kettle or the like is heated by hot fluids such as combustion gases supplied from an adjacent furnace

Apparatus for dispelling fog and mist with heated air. EDWARD C. READER U. S. 1,793,719, Feb 24 App is described suitable for use at aircraft landings, etc., which comprises an externally heated conduit through which air is forced which is then discharged as a forceful, voluminous blast into the fog or mist

Apparatus suitable for mixing measured portions of opaque substances with röntgenologists' solutions GEORGE W. HOPKINS U. S. 1,791,769, March 3 Structural features

Apparatus for "neutralizing" poisonous gases such as those from gasoline engines DANIEL A. MACKINNON U. S. 1,791,813, Feb 21 Structural details are described of an app. suitable for treatment of the gases with catalytic material such as oxidized Cu

Apparatus for electrical conductivity tests on water in bore holes (to determine character of surrounding rocks). Soc. de PROTECTION ELECTRIQUE (PROCEDES SCHLUMBERGER) Brit. 337,189, Oct. 26, 1928 Structural and elec features

Mercury-vapor boiler WALTER R. TARRANT (to Trent Process Corp.), U. S. 1,792,784, Feb 17 The interior of the boiler is surfaced with a material such as Ag which is wetted by Hg

Apparatus for testing wearing of gears, etc., in various atmospheres. M. FRANK Brit. 337,511, Aug 24, 1929 Wear is stated to be greater in gases contg O than in inert gas such as CO₂

Apparatus (with superposed mixing chambers and agitating devices) for dissolving xanthates RICHARD THUMM and EUGENE SCHMITZ (to Baker Perkins Co.) U. S. 1,791,214, Feb 24 Structural features

Apparatus for making amalgams, particularly for dental use. DRUTCHER GOLD-UND SILBER-SCHNEIDANSTALT VOM. ROESSLER Ger. 517,202, Mar 27, 1928.

Apparatus for estimating the humidity of foundry sand, etc., by determining the electrical resistance FERTZ FÖRST. Ger. 517,212, Dec 23, 1928

Valve for operation under high temperatures and pressures. W. PFRIFFER and K. RAUS Brit. 337,838, April 18, 1929

Acetylene generators PAUL WIDRACHSWA Swiss 142,901, July 22, 1929 App. for producing C₂H₂ from water and CaC₂ having a hollow vessel for regulating the C₂H₂ pressure, is described

Transport drum for carbide, etc. GUSTAV HILGER Ger. 517,281, Nov 10, 1929 Addn to 489,305 (C. A. 24, 2017)

Gas-cock control device FREDERICK A. GRACICH and WALTER M. BERRY (to Hammer Bray Co.) U. S. 1,792,488, Feb 17. Structural features

Automatic gas shut-off valve. LUTHER D. LOVEKIN (to Kilson Co.) U. S. 1,794,191, Feb 24 Structural details of a device which shuts off gas flow when water in a tank reaches a certain temp. Cf. C. A. 24, 2335

Thermostatic valve EDWARD S. HALSEY, U. S. 1,791,610, March 3 U. S. 1,791,611-12 also relate to valves of the same general type.

Thermostatic device suitable for control of electric circuits JOSEPH W. MYERS U. S. 1,793,934 Feb 24 Structural features

Thermostatic device for gas ranges. ARTHUR STOCKSTROM and ISAAC V. BAUMBAUGH (to American Stove Co.) U. S. 1,794,329, Feb 24

Thermostatic device suitable for humidity indicators, etc. ARTHUR P. KRICK (to National Dry Kiln Co.) U. S. 1,791,809, Feb 24 Structural features

Thermostatic valve for hot-water systems, etc. WALTER B. CLIFFORD U. S. 1,792,891, Feb 17

Thermostatic control device for burners such as those of water heaters. HOWARD I. PIATT (to Motor Wheel Corp.) U. S. 1,793,451, Feb 17. Structural features

Thermostatic control device for hot water and oil heat-storage systems. C. F. HAMMOND and W. SHACKLETON Brit. 337,425, July 31, 1929 Structural features

2—GENERAL AND PHYSICAL CHEMISTRY

FREDERICK L. BROWN

Adolf Claus. W. KLUGER, Brauer- und Hopfenzeitung-Gambrinus 58, 6-8 (1931), cf. C. A. 25, 1128—Obituary. S. LAUFER

Arthur Amos Noyes. MILES S. SHERRILL Ind Eng Chem 23, 443-5(1931)—Biography with portrait. G. G.

Hormann Reisenegger. H LAUBMANN *Chem.-Ztg* 55, 181(1931); *Z. angew. Chem.* 44, 213-6 —Obituary E. H.

To the memory of Karl Schirmacher. BRVK AND ZAHN *Z. angew. Chem* 44, 179-80(1931) —Obituary with portrait E. H.

Alfred Stovenhagen. H WÖHLING *Z. angew. Chem* 44, 178(1931) —Obituary with portrait E. H.

Edward William Voelcker. BERNARD DYER *Analyst* 56, 141-6(1931) —Obituary with portrait W T H

Nichols Medal award [to John Arthur Wilson]. The life of the medalist. CLARKE I. DAVIS *Ind Eng Chem* 23, 435(1931) Scientific accomplishments of the medalist. G C

ARTHUR W THOMAS *Ibid* 435-6 The history of The Department of Chemistry of The Ohio State University. WM. McPHERSON *J. Chem. Education* 8, 610-51(1931) E H

Chemistry in the vocational high schools of the middle west. MARLE SPENCER. *J. Chem. Education* 8, 712-6(1931) E H

The analytical distillation of gasoline. HAROLD J TORMEV *J. Chem. Education* 8, 679-705(1931) —A lab. exercise for students is described E. H.

The technic of micromanipulations. E A HAUSER *Rev. gen. colloides* 8, 358-61(1930), cf *C A* 22, 1405, 4476 25, 1409 —A lecture A P

Reflections on statistical thermodynamics. J H CONLYN *Chaleur et ind* 9, 465-76, 527-31(1929), 10, 191-9(1929) 11, 11 8, 141 6, 197-205(1930), cf *C A* 23, 2350, 24, 2661, 4984 S L B L. RHURTON

Simple reduction experiments. E H J MAGER *Z. physik. chem. Unterricht* 44, 25-6(1931) —M describes several demonstration expts. the reduction of boiling water with Zn dust, the reduction of CuO with (a) Fe dust, (b) wood charcoal in a porcelain crucible and (c) in a gas flame. CuCl₂ can be reduced to CuCl by the following treatment. Dissolve a small quantity of CuCl₂ in 3-4 cc. concd. HCl and place a wad of Cu wool in the soln. Boil the whole vigorously until the soln. becomes colorless. Pour a portion into water and a white ppt. of CuCl forms. cool the rest and shining crystals of CuCl appear. F R SCHMIDT

Gas-volumetric experiments in student exercises. JOSEF KIEMA *Z. physik. chem. Unterricht* 43, 265-7(1930) —K describes a simple and inexpensive app. for collecting gases which resembles the gas-collection part of the Victor Meyer mol.-wt. app. Two burets, one to serve as a leveling tube, are connected at their lower ends with rubber tubing. The one in which the gas is to be collected is fitted with a rubber stopper carrying a T-tube. One arm of the T tube is connected with a test tube or side-neck test tube which serves as a reaction chamber, the other arm is closed with a pinch clamp. This app. with slight modifications to suit individual cases can be used to test the reactivity of metals, the amt. of O in HgO, and for the quant. reduction of oxides (HgO, CuO) or the oxidation of Cu. Diagrams and directions are given. F R S

Evolution of hydrogen sulfide in the Bay of Krasnovodsk. D ZANKHIV *Azerbaidzhanskoe Neftyanoe Khozaystvo* 1930, No. 10, 58-61 —The role of microorganisms is described V KALICHEVSKY

Preparation of pure mercury. H VON STEINWEHR AND A SCHULZ *Z. Instrumenten* 49, 566-7(1929) —Hg for use in phys. app. is repeatedly shaken with dil. HNO₃ and then with water, fat is removed by means of petroleum or ale. The Hg is then filtered through a thin leather, previously freed from fat with petroleum, or a Haldenwanger filter tube. It is then distilled in a current of air in a water pump vacuum in an app. without taps, and preserved in evacuated glass vessels. B C A

Final report of L. M. M. representatives on the Sieves and Screen Committee of the British Engineering Standards Association. S J TRUSCOTT AND H I. SUTMAN *Bull. Inst. Mining Met.* 311, 9 pp (1930) —The proposed fine testing sieves are based on a ratio of the fourth root of 2 between linear aperture measurements of consecutive sieves. The American and proposed British sieves will be essentially identical as to aperture but not as to mesh because of the difference in American and British wire gage. ALDEN H. EMERY

A note on gravity separation. R. C. EMMONS *Am. Mineral* 15, 536(1930) —The unsatisfactory nature of gravity sepn. by means of heavy solns. on material in a fine state of division is best accounted for by adhesion, both of grain to grain and grains to air bubbles. This can be eliminated by strong evacuation of the sepn. vessel. A few min. is sufficient. A rapid stratification results when the liquid is allowed to stand, not necessarily under vacuum. A. M. BRANT

The technical preparation of divi-manganese. J. G. F. DRUCK *Ind. Chemist* 7, 75, 79(1931), cf *C A* 24, 2352 —The history of divi-Mn (Re), at no 75, since its dis-

the principal dipole. The moment of the iso-compd is the same as that of the normal if the branch in the chain is at least 2 C atoms away from the halogen, because only these 2 C atoms are affected by the principal dipole.

H W LEAHY

Superstructure and magnetic susceptibility in the system: copper-gold. H J. SEEMANN AND E VOGT *Ann Physik* [5], 2, 976-90(1929), cf *C A* 24, 5196.—The variation of diamagnetic susceptibility with concn was measured for the Au-Cu mixed-crystal series. Alloys of the compns Cu_3Au and CuAu in certain circumstances exhibit superimposed structure lines consistent with a regular arrangement of the constituent atoms, i.e., the corners of the elementary cube may be occupied by Au atoms and the face centers by Cu atoms. The relation between susceptibility and the appearance of superstructure was detd for alloys of the compns Cu_3Au and CuAu . For the former an increase and for the latter a decrease in diamagnetic susceptibility accompany the production of regular distribution in the lattice. Control expts on alloys differing in compn from the whole no ratios indicated no difference in susceptibility for quenched and slowly cooled specimens. No explanation is available for the difference in sign of the effect with Cu_3Au and CuAu .

B C A

Magnetic measurements on rhenium. W H ALBRECHT AND E WEDEKIND *Naturwissenschaften* 19, 20-1(1931).—Samples of Re and of K perberrate having less than 0.01% impurity in the form of Ca and U were measured. Re gave at 18° a sp susceptibility $\chi = +0.046 \times 10^{-4}$ (≈ 0.02), the perberrate is diamagnetic $\chi = -0.13$ (≈ 0.05). The Re value does not lie between those of the neighboring elements Os ($+0.074 \times 10^{-4}$) and W ($+0.33 \times 10^{-4}$).

B J C VAN DER HOEVEN

Relationships between the dielectric constant and the optical properties of substances having the sodium chloride lattice. J ERRERA *Z Elektrochem* 36, 818-23 (1930).—The total polarization of the alkali metal halides increases with increase of at wt of the anion and the cation. By using Spangenberg's values for the electronic polarization, the at polarization was calcd. The proportion of the total dielec. const due to at vibrations is decreased as the at no of the anion is increased, and in the case of the bromides and iodides with increase of at no of the cation. There is an irregular variation in the chlorides and fluorides. The factor p of the infra-red term of the dispersion formula can be detd by comparing the characteristic infra-red frequencies calcd. by the dispersion formula from the observed values of the dielec. const. with those calcd. by Born by the theory of electrostatic cohesion. The infra-red branch of the dispersion curve can also be calcd.

L L QUILL

Action of nitric acid on benzoic acid in magnetic and in electrostatic fields. R. H. CLARK AND R. M. ARCHIBALD *Trans Roy Soc Can* [3], 24, Sect 3, 121-4(1930); cf following abstr.—The action of HNO_3 on benzoic acid was carried out at 20° in a magnetic field of 23,000 gauss per $\pi/4$ sq cm. The products of the reaction did not vary appreciably from those obtained outside the magnetic field. J W SHIPLEY

Addition of hydrogen bromide to allyl bromide in a magnetic and an electrostatic field. R. H. CLARK AND K. R. GRAY *Trans Roy Soc Can* [3], 24, Sect 3, 111-4 (1930).—The effect of a magnetic field on the reaction of HBr and allyl bromide in glacial AcOH decreases the yield of trimethylene bromide. For propylene bromide $n_D^{20} = 1.51904$ and for trimethylene bromide $n_D^{20} = 1.52292$.

J W SHIPLEY

Further investigation of the two electromers of 2-pentene. R. H. CLARK AND E. G. HALLONQUIST *Trans Roy Soc Can* [3], 24, Sect 3, 115-9(1930), cf preceding abstr.—The 2 electromers of 2-pentene were synthesized and some of their phys. properties detd. Neither a magnetic nor an electrostatic field had any effect upon the addn products obtained with HBr on the metastable 2-pentene. Stable 2-pentene has the following consts: b p 35.8° (755 mm), n_D^{20} 1.37845, f p -149°, metastable 2-pentene has b p 36.5°, n_D^{20} 1.37960, f p -141°.

J W SHIPLEY

Sound propagation in gas mixtures. D. G. BOURGIN *Phys Rev* [2], 34, 521-6 (1929).—Mainly math. Previous results on a single gas and a mixt. of 2 gases (cf *Phil Mag* [VII], 7, 821-41(1929)) are simplified and extended to a mixt. of any no of gases.

B C A

Calculations on the velocity of sound in nitrogen tetroxide. FRANK VERHOEK AND FARRINGTON DANIELS. *J. Am. Chem. Soc.* 53, 1186-7(1931).—By using dissociation data of N_2O_4 detd. by V. and D., the velocity of sound was calcd. and compared with the velocities detd. experimentally by Kistiakowsky and Richards (*C. A.* 25, 864). The agreement is not much better than with the data of Bodenstein, used by K. and R., but the deviation is nearly const. and the calcd. value is always less than the observed velocity, this may indicate a const. error, either in theory or expt.

C. J. WEST

The formation of heams in molecular streaming. P. CLAUSING *Z Physik* 66,

471-6(1930), cf *C. A.* 24, 3403—Formulas are given for the distribution of moles that stream through a short tube at const. pressure into an evacuated space. The mean free path of the moles is great compared with the size of the tube and diffuse scattering from the walls of the tube is assumed. In the case in which the length of the tube equals the diam., large deviations from the cosine law occur. G M M

Viscosity, heat conduction and diffusion in gas mixtures. XL The viscosities of H_2 , N_2 , CO , C_2H_4 , O_2 and their binary mixtures. MAX TRAUTZ AND ALBERT MELSTER *Ann. Physik* [5], 7, 409-26(1930), cf *C. A.* 25, 10—The viscosities of H_2 , N_2 , CO , C_2H_4 , O_2 and their binary mixts. were detd. over the temp. range 20° to 250° by measuring the rate of flow through a capillary tube. The results were well represented by the Sutherland formula, $\eta/\eta_0 = (T_0/T)^{1/2} \times [1 + (C/T_0)]/[1 + (C/T)]$ by using the following values for the const. C : H_2 , 84.4, N_2 , 103.9, CO , 101.2, O_2 , 126.6, C_2H_4 , 241.4. The isotherms of N_2 - CO , N_2 - O_2 and CO - O_2 mixts. were practically straight lines. The relation between viscosity and mol. diam. in mixts. is discussed. XII. Viscosity of gases at high temperatures. MAX TRAUTZ AND ROBERT ZINK. *Ibid.* 427-52—The viscosities of air, N_2 , H_2 , O_2 , CO_2 , SO_2 , CH_4 , A , Ne and He were measured over the temp. range 20° to 830° . The viscosities of air, N_2 , A , O_2 , CH_4 , SO_2 and CO_2 , as well as of HCl and Br_2 , are accurately represented by the Sutherland formula. The values of C for the different gases are given. This formula fails for H_2 , He and Ne below the crit. temps., but the viscosities of these gases are given by the formula $\eta = kT^n/[1 + (C/T)]$ where n and C have the following values: H_2 , $n = 3/4$, $C = 0$, He , $n = 1/2$, $C = 2.5$, Ne , $n = 0.6$, $C = 20$. P T NEWSOME

Calculating viscosity and flash point in compounded oils. E KADNER *Seifen- und Zieg.* 56, 22-3(1931)—See *C. A.* 25, 805. P FORTNER

Note on the equation of state explicit in the volume. GEORGE SCATCHARD *Proc. Natl. Acad. Sci.* 16, 811-3(1930)—The Beattie-Bridgeman equation (*C. A.* 21, 3146) may be expressed in the form $V/n = (RT/p) + (b/RT) + [\gamma/(RT)^2 - \beta/(RT)]p + [1/(RT)^2 - 3\beta\gamma/(RT)^4 + 2\beta^2/(RT)^6]p^2$, in which n is the no. of moles and the coeffs. β , γ , δ are functions of the temp. This equation reproduces the explt. results for less perfect gases less satisfactorily than Beattie's original equation (*C. A.* 24, 1775). The equation is not accurate at high pressures. J B AUSTIN

The specific heat of liquid biphenyl. ROY P. NEWTON, B. D. KAURA AND THOMAS DU VRIES. *Ind. Eng. Chem.* 23, 35-7(1931)—The sp. heat of liquid Ph_2 up to 300° was detd. by 2 elec. heating methods and by a method of mixts. The 3 methods gave consistent results with an av. value for the sp. heat of $(0.338 + 0.00057t) = 1^\circ C$. F. H. RATHMANN

The specific heat of biphenyl. H O FORREST, E W BRUGMANN AND L W T CUMMINGS. *Ind. Eng. Chem.* 23, 37-9(1931)—The sp. heat of Ph_2 in small temp. intervals from 72° to 203° was detd. by the batch-calorimeter method and from 147.7° to 347° by the flow calorimeter. The sp. heat varies from $0.403 = 2^\circ C$ at 77.6° to $0.672 = 2^\circ C$ at 347° . The results differ from those of N., K. and De V. by from -5 to $+12\%$ (cf preceding abstr.). F. H. RATHMANN

The necessity of adopting a standard substance for boiling-point and vapor-tension measurements. W SWIETOSLAWSKI *J. chim. phys.* 27, 496-502(1930), cf following abstr.—After a discussion of work done by others in this field, it is recommended by the Comm. on Physical-Chemical Standards of the International Union of Chemistry that b. p. and vapor tension be detd. simultaneously with the b. p. and vapor tension of water, that a few other liquids be standardized and that dp/dt be detd. G M P

A comparative study of boiling points and vapor tensions of eight standard organic substances. AL ZMACZYNSKI *J. chim. phys.* 27, 503-17(1930)—B. ps. and vapor pressures of 8 compds. were measured between 380 and 2000 mm. pressure. The b. ps. of the substances studied were compared with the b. p. of water by the formula $t_{\text{sub}} = A + Bt_{H_2O} + Ct_{H_2O}^2$, and the coeffs. A , B and C were calcd. dp/dt was calcd. The b. ps. detd. at 760 mm. were: $EtBr$, 33.386° ; CS_2 , 46.262° ; Me_2CO , 56.131° ; $CHCl_3$, 61.152° ; C_6H_6 , 80.122° ; $PhNe$, 110.606° ; $PhCl$, 131.687° ; $PhBr$, 155.908° . GERALD M. PETTY

The vapor pressure of pyridine. P A VAN DER MEULEN AND RUSSELL F. MANN *J. Am. Chem. Soc.* 53, 451-3(1931)—The vapor pressure of pyridine was detd. by the isoteniscope method. Between -20° and $+120^\circ$ the pressure is given by the formula $\log p = 6.8827 - \frac{1281.3}{t + 205}$, where p is expressed in mm. of Hg and t in degrees C. P T NEWSOME

Systems of four immiscible liquid layers. E LESTER SMITH *Nature* 127, 91 (1931)—An example of the rarely observed phenomenon of 4 immiscible liquid

layers can be produced by using hexane 12 cc, PhNH₂ 7 cc, oleic acid 0.5 cc, alc 1.5 cc and 0.8 N NaOH 10 cc. The alc is not essential to the system but reduces the time required for sepn of the layers. An increase of temp or of electrolyte renders miscible the hexane and PhNH₂ layers (first and third from the top), which then form the top layer. H. P. JOHNSTON.

Relation of the liquid to the crystalline state. KENARISWAR HANTRIJE. *Natura* 127, 92-3 (1931).—The origin of elastic heat waves in liquids is elastic and akin to that of lattice oscillations in crystals. From theory and from observations on x-ray diffraction, it is seen that there is a max in the mol arrangements in liquids corresponding to the structure of the substance in the crystal state. In a crystal the natural frequencies of a mol that are responsible for the Raman effect are modified because of the surrounding mols, but on account of the regular arrangement all the mols are affected in the same manner and the Raman lines are sharp. In a liquid, however, the regularity of arrangement is broken, and there is a broadening of the natural frequencies of the mols. H. P. JOHNSTON.

Loading of metals with gas at high pressure. HANS JELLINEK. *Z. Physik* 66, 547-57 (1930). After activation, i.e., Pd and Ag take up H₂ and N₂ under high pressures. The phenomenon is studied by means of measurements of the elec cond. When the gas is taken up by the metal, the resistance increases and falls again when the pressure is removed and the gas is liberated. The resistance for Pd and Ag increases 40% between 0 and 35 atm, but for Fe and Ni it decreases 45% up to 15 atm and then increases 17.3% from 15 to 30 atm. The increase in resistance and pressure range for other cases are: Pd-H₂, 4.1%, 0-20 atm; Pd-N₂, 3.8%, 0-25 atm; Ag-H₂, 13.1%, 0-35 atm. G. M. MURPHY.

Measurements with the help of liquid helium. XII. Plasticity of metal crystals at low temperatures. W. MEYER, M. POLANYI and I. SCHMIDT. *Z. Physik* 66, 177-89 (1930); cf. *C. A.* 24, 3087.—The plastic deformation of Zn and Cu crystals on stretching was investigated over the temp range 1-20° abs. The plasticity remains nearly const. A method for transporting and handling liquid He is described. G. M. M.

Dependence of the dissociation temperature of solids on the size of crystal grains. J. KRUSTINSON. *Acta Univ. Lundsens. Kim. Fakult.* 1, 273-7 (in German 278) (1929); cf. *C. A.* 23, 417.—Reil HgO was sepd. into 4 fractions by sieves and it is shown that the dissoci. temps. of the smallest and largest sizes differ by 6°. If the Thomson formula is assumed to hold, the calcd. surface tension at the gas-solid interface is 1110 dynes/cm. G. M. MURPHY.

The conductivity of thin metal films. LADISLAUS TISZA. *Naturwissenschaften* 19, 60-7 (1931).—The limiting thickness of thin metal films (by cathodic dispersion) at which their resistance rises sharply (cf. Bartlett, *C. A.* 22, 2470; Bernece Ann. *Physik* 4, 252 (1930); Braunbek, *C. A.* 24, 1657) can be derived from the free path of the cond. electrons as indicated by J. J. Thomson. It is assumed that the electrons are diffusely reflected by the surface and thereby lose their forward impulse. The resulting free path is $\lambda' = \lambda d / (d + \alpha)$, in which λ is the free path in compact metal, d the thickness of the film, α a const. of about $1/2$. The crit. limiting thickness is $d_c = \text{about } \alpha \lambda$, which agrees with exptl. values (10 μ for Ag at room temp., $\lambda = 60 \mu$, Sommerfeld). At low temps. with d_c many times larger (λ longer) the results should be more pronounced and characteristic. The present theory, however, does not account for Braunbek's result on Hg: its d_c value of 200-300 μ at -153° is too high. H. J. C. V. D. H.

Polar properties of single crystals of ice. J. M. ANASTAS. *Proc. Roy. Soc. (London)* A128, 688-91 (1930).—Photographs are reproduced of single microscopic ice crystals grown under controlled conditions. A short right hexagonal prism, with the height nearly equal to the diam. of the base, is a frequently occurring shape for these small crystals. A no. of specimens give evidence of crystallographic polarity by the appearance of a jut at one end only of the c axis when the crystal is brought into an atm. favorable to evapn. Two other types of disintegration, characterized resp. by juts at each end and in the middle of the c axis, are observed in crystals of the same external form. These 2 types are explained as being due to the 2 possible modes of twinning of a polar crystal on the basal plane. The conclusion that the c axis in ice is a polar axis does not necessarily contradict the conclusion reached by Barnes that the (0001) plane must be a plane of symmetry so far as the arrangement of O ions in the lattice is concerned (*C. A.* 24, 765). It is possible that the asymmetric location of H ions in a non polar O lattice confers the necessary polarity on the structure as a whole. B. C. A.

Crystal structure of molybdenum trioxide. NORMA WOOTER. *Natura* 127, 93 (1931).—Lane photographs of crystals of MoO₃ show orthorhombic symmetry. The cell size is $a = 3.97 = 0.02 \text{ A. U.}$, $b = 13.01 = 0.05 \text{ A. U.}$, $c = 3.57 = 0.02 \text{ A. U.}$ There

are 4 mols of MoO_3 per cell. The space group is Q^{14} . The crystals were thin plates parallel to (010). H. F. JONSSON.

The grating constant of quartz. O. BERGQVIST. *Z. Physik* 56, 494-8 (1930).—The grating const. of a rock crystal was measured in the 1st and 2nd orders with Cu K_{α} and $1.6 K_{\alpha}$. The distance from slit to plate was 622.11 mm; the slit width, 0.62 mm. The x ray tube operated at 35 kv and 5 milliamperes, the exposure time was from 20 min. to 2 hrs. No measurable plates were obtained in the 3rd order in 2 hrs. From the values in the 1st order, the higher-order results were calculated. A correction curve to the wave-length calcul. is given for the anomalous dispersion of quartz. The results are $d = 4246.92 \text{ \AA}$, $\log 2d = 3.9240120$. G. M. MURPHY.

Crystal structure of nickel oxides. O. G. BENNETT, R. W. CARRAS, and EMIL OTT. *J. Am. Chem. Soc.* 53, 1179-10 (1931).— NiO , prep'd by oxidation of $\text{Ni}(\text{OH})_2$ with Cl in alk. soln. and heated at 285° , correspond to the previously reported structure for NiO , as it is face centered cubic. If the NiO is not heated above 110° , the comp'd. is cubic but the edge of the unit cube is 4.61 \AA , the d is 4.8 , as compared to 6.69 for ordinary NiO ; the no. of mols in the unit cell is calcd. as 4; consideration of the intensities of reflection excludes the face-centered cubic type. There is evidence of further modifications. C. J. WEST.

The physical properties of rhodium. C. AGTE, H. ALTHERTUM, K. BECKER, G. HEYNE, and K. MOHR. *Naturwissenschaften* 19, 104-9 (1931).—Several properties of Re were det'd. The crystal structure is hexagonal in close packing, const. $a = 3.765 \text{ \AA}$, $c = 4.470 \text{ \AA}$, $c/a = 1.16$, smallest distance 2.765 \AA , atomic radius 1.382 \AA . The μ is 3440 abs. (measured in drilled hole), the linear thermal expansion (x ray method) is 12.45×10^{-6} along the hexagonal axis, 4.67×10^{-6} perpendicular to it (5% accurate). The α of a highly annealed rod is 20.9 (calcd. 21.4 for wt. $180.31 = 0.02$). The fracture strength of a Re wire grown in gas phase, 0.25 mm diam., 0.03 mm W core in 23.6 kg per sq. mm, elongation 21% . The sp. elec. resistance at room temp. is $\rho = 0.21 \times 10^{-8}$ ohm cm (15% accurate), its temp. coeff. between 0° and 100° is 3.11×10^{-4} , between 0° and 2710° , 1.93×10^{-4} . The ratio of ρ_{2710}/ρ_{100} is 6.31 . The electron emission between 1900° and 2700° abs. gives for the Richardson equation $A = 290 \text{ amp. per sq. cm}$ and degree squared, $b = 50,500$, corresponding to an exit energy of 5.1 eV . B. J. C. VAN DER HOEVE.

Condition magnitudes for water in the range from 10° to 500° and from saturation pressure to 300 atmospheres. M. TRAUTE and H. STIEHA. *Forschung Gebiete Ingenieurw.* 2, 45-51 (1931).—A high pressure vessel (vol. 1410 cc. at 15° and 1 atm) was fitted with a heating device, thermoclements and manometer, and placed in a container packed with insulating material. An air space sep'd this container from the outer jacket. The sp. vol. of water was measured in the range 10° to 370° and 150 to 300 atm . Below the crit. temp. the std. accuracy was about $\pm 0.2-0.4\%$, and at the crit. temp. $\pm 2.4\%$. The change in heat content between 10° and 500° , over the range 150 to 300 atm , was measured with an std. accuracy of $\pm 3\%$. From the present PVT data, together with values interpolated between 150 atm and satn. pressure by utilizing the International Skeletal Steam Tables, a series of values were calcd. for (1) c_p , the product of the Joule-Thomson effect and the sp. heat, (2) J , the heat content, and (3) C_v , the sp. heat at const. pressure. F. D. ROSSINI.

Thermal properties of compressed liquid water. JOSEPH H. KEENAN. *Mech. Eng.* 53, 127-31 (1931).—Recent data of Keyes and Smith on compressed liquid sp. vol. measurements and of Osborne, Stinson and Fock on sat'd. vapor measurements, have made possible a much more accurate calcul. of the feed pump work in power plant operation in the higher temp. and pressure ranges. Calcul. are based on the more accurate const. entropy measurements as compared with the older and less accurate const. vol. method. Temp. entropy diagrams are given for ranges up to 700°F . A diagram is given showing a comparison between const. entropy feed pump work and const. vol. feed pump work for abs. pressures from 0 to 400 kg./cm^2 . The differences are small, usually within 2% of the compression work. A. L. KILMER.

Some final values for the properties of saturated and superheated water. FREDERICK G. KEYES and LEONARD B. SMITH. *Mech. Eng.* 53, 132-5 (1931).—Tabulations are given of all available data of the Mass. Inst. of Technology pertaining to the properties of water. The tabulations include the following: vapor pressures between 100° and 374° , dp/dT at satn. between 0° and 373° , sp. vols. of sat'd. steam between 100° and 370° , sp. vols. of water at satn. pressure between 100° and 374° , d of water between 100° and 370° , vols. of liquid water under pressure and at different temps.; isometrics of superheated steam between 200° and 400° . The use of an alloy of $18\text{Cr}-8\text{Ni}$ steel increased the accuracy of high temp. measurements. A. L. KILMER.

Some additional volume data for superheated steam. L. B. SMITH AND P. G. KRYES *Mech Eng* 53, 175-7 (1931)—The temps at which steam has been measured are extended to 460° by the use of a stainless steel bomb. Results are given in tables and diagrams.

A. L. KILPER

The vapor pressure of chlorine monoxide. CHARLES F. GOODEVE. *J. Chem. Soc.* 1930, 2733-7—The vapor pressures (from 288° abs to 173° abs), the m p and the b p of ClO are detd. The plot of log vapor pressures (mm) against $1/T_{\text{abs}}$ is the straight line $\log_{10} p = -1773/T + 7.87$. From the vapor pressures the b. p. of ClO is 20°/760 mm, which is a better value than that of Goldschmidt (*C. A.* 14, 20). From the slope of the line the latent heat of evap. (L_v) is 6200 cal/g-mol. The m p of ClO is detd from cooling curves as $-116 \pm 1^\circ$. The value of Trouton's coeff indicates that little or no assocn of ClO takes place in the liquid state.

J. BALOZIAN

The electrical conductivity of amorphous quartz. WALTER GNANN. *Z. Physik* 66, 430-52 (1930), cf *C. A.* 24, 5197—The surface cond of amorphous quartz cylinders of 1 cm diam and 1.2 cm length is measured as a function of humidity of the air, tension and time. The cond is strongly dependent on humidity. Up to 20% relative humidity, quartz is a better insulator than amber, but its cond increases more rapidly with humidity. Immediately after tension is applied, the current is relatively large but drops after a few hrs to a const value. The vol resistance is greater than 6×10^{10} ohm cm. Similar measurements made on Bakelite C show that it is strongly hygroscopic.

G. M. MURPHY

Blanc's alumina. N. PARRAVANO. *Mem. accad. Italia, Classe sci. fis. mat. e nat. I, Chim.* No. 1, 1-27 (1930), cf *C. A.* 24, 4681—Further study of Blanc's Al_2O_3 reveals powerful adsorbing properties. At 25° it retains 32.5% CCl_4 , 18.7% C_2H_6 , 18.2% CS_2 and 17.9% Me_2CO . It is anisotropic; a diffraction spectrum appears only after mol rearrangement, which occurs at 180° to 670°, depending on exptl conditions. This change is accompanied by a crowding of the atoms in the space lattice, which is probably the source of the heat developed in the process. X-ray analysis also shows that this second form passes into a third, corundum, above 970°, this reaction is also an exothermic one accompanied by further crowding of atoms in the lattice. This opinion is strengthened by the fact that the corundum form, is not a catalyst because it will not dehydrate H_2OII , while the other forms, whose atoms are further apart, have strong catalytic powers. A summary of the previous work on this substance is included.

A. W. CONTIERI

Gradual transition in crystalline sodium nitrate. P. C. KRACK AND E. POSNJAK. *J. Am. Chem. Soc.* 53, 1183-4 (1931)— NaNO_3 exhibits a gradual transition which ends at approx. 275°; the properties of the crystals alter reproducibly over a range of temps. rather than suddenly at a definite transition pt. The coeff of expansion is nearly const. to about 160°, then it gradually increases to a peak value near 275°, followed by a rapid decrease to a normal value beyond 280°. The heat absorption on heating likewise gradually increases to a max at 275.5°, while on cooling the heat evolution begins at about 275°. Other changes are noted.

C. J. WEST

Calorimetric determinations of heats of adsorption. Adsorption of sulfur dioxide by wood charcoal. A. MAGNUS, H. GIPFENHAIN AND H. VELDP. *Z. physik. Chem.*, Abt. A, 150, 285-94 (1930)—The integral heats of adsorption of SO_2 by coconut charcoal were measured over a wide range of concn (0.05-64 mg of SO_2 per g of charcoal). With increase in the amt adsorbed there is a slow, regular fall of molar heat of adsorption. Differential heats of adsorption were measured by allowing about 0.12 mg of gas per g of adsorbent to be adsorbed on the already more or less charged charcoal. A marked discrepancy appears between the curves for integral and differential heats of adsorption, and this is explained by the slow establishment of equil in the crevices on the surface of the adsorbent.

B. C. A.

Adsorption influence, activity and solvation in salt solutions. P. P. KOSAKOVICH AND N. A. ISMAILOV. *Z. physik. Chem.*, Abt. A, 150, 295-309 (1930)—The adsorption by charcoal of PhOH in aq soln and of H_2OII in MeOH or EtOH soln is increased by the presence of the chlorides or bromides of Li, Na or K. In aq soln the relative effects of the cations are in the order $\text{Na} > \text{Li} > \text{K}$ and those of the anions $\text{Cl} > \text{Br}$. In MeOH the same order is observed for the anions, but this order is reversed in EtOH . The salts raise the activity of the substance adsorbed, providing a means of calcg the relative activities of the latter from adsorption measurements. On the assumption that the neutral salt effect is due to solvation, values are obtained from adsorption measurements in good agreement with those calcd. from the influence on soly.

B. C. A.

HCHO breaks emulsions stabilized by proteins (4) Anhyd Na_2SO_4 , $(\text{NH}_4)_2\text{SO}_4$, or CaCl_2 often salts out emulsifying agents (5) Filtration breaks emulsions stabilized by flocculent ppts W. G. GAESSLER

The dielectric constants of complex colloidal systems. Adsorption by micelles in solution. Dielectric origin of the adsorptive forces. CHARLES MARIE AND NÉDA MARINESCO *J chim phys* 27, 455-70(1930), cf *C A* 21, 3806, 22, 4385—Tables and graphs are given which show the variation of dielec const with concn of suspensions of C in water and in gelatin, of isolec gelatin, of l leucine and C, and of glycine, gelatin and C The phenomenon of protection was sepd from the phenomenon of assocn in complex mixts Dielec analysis permits the arbitrary choice of a protective colloid, and a detn. of the quantity necessary to protect an unstable colloid

GERALD M. PETTY

Formation of spirals by chemical precipitation. R. E. LIESEGANG *Naturwissenschaften* 18, 645-6(1930)—The spiral structures found by Schikorr (*Ibid* 18, 376(1930)) are simply varieties of the ordinary Laeegang rings A photograph showing the occurrence of these spirals in Ag_2CrO_4 pptd in gelatin is reproduced B. C. A.

Study of E. Fischer's polypeptides with x-rays. FRITZ V. LENEL *Naturwissenschaften* 19, 19(1931)—X-ray photographs were made of 4 Fischer preps I dl leucyltetraglycylglycine, II l leucyloctaglycylglycine, III l leucyltriglycyl l leucyloctaglycylglycine and IV l leucyltriglycyl l leucyltriglycyl l leucyloctaglycylglycine "Mark" tubes of 0.8 mm diam., 10 to 20 μ wall were used to contain the substances in a Debye-Scherrer camera of 117.7 mm diam., Cu K rays, 10 hrs exposure, 10 to 12 amps, 35 kv Rings at various angles were found In the last 3 preps a recurring ring at $\theta = 10^\circ 36'$ indicates a crystallographic distance of 4.07 A U between polypeptide chains (4.65 and 5.20 for silk) For I a distance of 3.15 A U is found that is probably unrelated to the 3.5 A U of the av. of glycylalanyl (K. H. Meyer *C. A.* 23, 3721) A pos result of the study is that the polypeptides are cryst B. J. C. VAN DER HORST

Solubility of salts in ethylene glycol and in its mixtures with water. H. M. TRIMBLE *Ind Eng Chem* 23, 165-7(1931)—Tabular and graphical data are presented for the soly at 30° of KI , KBr , CuSO_4 , 5 aq, K_2SO_4 , NaCl and KCl in ethylene glycol, in water and in mixts of the 2 solvents The relative order of solubilities is much the same Metathetical reactions, and reactions in which one metal displaces another, took place readily in ethylene glycol solns

GERALD M. PETTY

Influence acting upon solubility. VII The couples: thiourea-antipyrine and caffeine-antipyrine E. OLIVERI-MANDALA AND L. IRRERA *Gazz chim ital* 60, 872-7(1930), cf O-M, *C A* 21, 1395—The cryoscopic measurements were carried out by the method of Bargellini (*C A* 14, 488) The expts show that (1) the solubilities of thiourea (I) and of caffeine (II) in water increase greatly with addn. of antipyrine (III), (2) the temp coeffs of soly and the thermal effect, q , calcd by the isochore formula of van't Hoff are not reliable criteria of the presence of mol. complexes in water; and (3) the cryoscopic method and that based on soly coeffs at different temps often give different results The following data give the soly (mols. per l) in water of I and of II at the temps shown: I; 10° , 1.1176, 15° , 1.383, 20° , 1.544, 25° , 1.759, II, 15° , 0.058, 25° , 0.100, 35° , 0.167 The soly values and soly coeffs of the I-III and II-III solns are given in more extensive tables, which should be consulted for details

C. C. DAVIS

Critical solution temperatures of systems of sulfur dioxide and normal paraffins W. F. SEYER AND E. TODD *Ind Eng Chem* 23, 325-7(1931)—It was found in study of the systems of SO_2 and butane, hexane, octane, decane, dodecane, tetradecane and tetradecone that the crit soln temp is a function of the mol wt of the hydrocarbon The amt of hydrocarbon sol in SO_2 at its b p is comparatively small

D. F. BROWN

Temperature of maximum density of aqueous solutions. Deviations from the law of Despretz. NORA GREGG-WILSON AND ROBERT WRIGHT *J. Phys Chem* 35, 624-8(1931), cf *C A* 13, 923—Previous expts with dil solns supported the law of Despretz, but on investigating more concd solns deviations were found which seem to depend upon 3 factors, viz., lowering of the f p, increase in the coeff of expansion and the d of the soln In general with more concd solns the temp of max d is lower than the f p The law of Despretz holds only for dil solns H. W. LEAHY

The role of chemical impurities in the fluctuation of heats of solution of monotropic salts. K. P. MISHCHENKO *Z Elektrochem* 36, 777-82(1930)—In order to det. the source of the variations in the heat of soln of KCl (of the order of ≈ 1 percent) very accurate measurements were made with an adiabatic calorimeter (cf *C A* 24, 4694) devised by L. and M. Other solid modifications of KCl were considered as a possible

source of the variations and for this reason certified samples of KCl from two sources were used after drying at 300° , after fusing, and with no previous treatment. Excellent agreement at 25° of 6 detns. with the high—4200 cal and the low—4104 cal., gave an av. value of —4196 cal as the integral heat of soln. of KCl to KCl 200 H₂O and indicated that other modifications were not the source of error. A sample of KCl of analytical grade gave consistently low results—about —4175 cal but after one recrystn it gave —4184 cal. This indicated the source of error and when addn. of small amts. of CaSO₄ to the certified samples gave correspondingly low results the proof seems conclusive.

ODEN E. SHEPPARD

Heats of dilution of uni- and multivalent strong electrolytes at great dilution. E. LANGE AND J. MONHEIM. *Z. Elektrochem.* 36, 772 (1930).—The exptl. integral heats of diln. recently made available by Lange and collaborators (cf. *C. A.* 22, 346, 24, 4004, 23, 1562 etc.) for 1-1 type, 2-2 type and 4-2-1 type salts are collected, plotted, and discussed in a more extended fashion (cf. Lange and Meixner *C. A.* 24, 1015). With 1-1 type salts with same cations the mutual exptl. slopes, Δ , which correspond to a'' values (apparent ionic diameters) are larger the smaller the radius of the opposite ion. This relation is valid for the 2-1 type salts investigated. $\Delta_{KCl} > \Delta_{CaCl_2} > \Delta_{MgCl_2} > \Delta_{NaCl}$. However the measurements available indicate the reverse condition in regard to the 2-2 type salts as far as is indicated by CaSO₄ and MgSO₄. Possibly this is due to the fact that Mg ion has a greater parameter due to the extent of hydration. Possibly in all cases the temp. change of the parameter is a factor individual for each ion. This may also play a part in the explanation of neg. slopes in heats-of-diln. curves.

ODEN E. SHEPPARD

Rotation dispersion of optically active substances. F. KERSER. *Strahlentherap.* 34, 574-7 (1929).—The circular dichroism and optical activity of various proteins were studied; the curve of the anomalous rotation dispersion of a 5% egg-albumin soln. in presence of Cu sulfate is reproduced.

H. C. A.

The theories of the Soret effect. G. S. HARTLEY. *Trans. Faraday Soc.* 27, Pt. I, 1-10 (1931).—H. reviews the kinetic theories of the Soret effect advanced by van't Hoff, Wrede, Porter and Chapman and points out the error common to all of them, viz. attributing to the solvent only the role of vehicle and carrier of the solute. He obtains independent support for Eastman's theory by basing a treatment on the assumption that a gradient of osmotic pressure produces the same velocity of ion transfer as an elec. field which gives rise to the same force per ion.

F. R. SCHERER

Note on the use of acetamide as a solvent in cryoscopy. E. CHIFFRUTIERZ AND G. DE MANDROT. *Helv. Chim. Acta* 14, 183-6 (1931), cf. *C. A.* 25, 2159.—Acetamide (I) was the only suitable solvent found for detg. the mol. wt. of certain disintegration products of casein. Previously detd. (cf. Bruni and A. Manuelli, *Gazz. chim. ital.* 35, 1418 (1905)) values of the cryoscopic K for I (3630 with urethan, 3130 with naphthalene and 3740 with hippuric acid) disagreed so widely that new values were detd. with acetamide, where $K = 4861$ at 0.897% concn., 4223 at 1.877%, and 4177 at 2.520%; hippuric acid, where $K = 4543$ at 1.346% concn., 4210 at 2.418%, and diketopiperazine, where $K = 4885$ at 0.507% concn. and 4790 at 0.460%. The values of K decrease with increase in concn. but at the same mol. concn. of each solute the values check within 5%. The detns. were made with pure, dry I in Beckmann tubes in a paraffin bath (to avoid water vapor) placed in a const.-temp. enclosure at 78° maintained by boiling AcOEt. I kept in a desiccator does not melt at a const. temp. and in a week the m. p. drops several degrees, indicating 2 modifications of I. Despite these facts results agree well when values of K are used depending on the concn., and the results are considered valuable in the work under consideration.

ODEN E. SHEPPARD

Solvents. T. H. DERRANS. *J. Oil & Colour Chem. Assoc.* 14, 65-80 (1931).—The phys. chemistry of solns. with particular reference to lacquer solvents is discussed.

G. G. SWARD

Studies on the porous-disk method of measuring osmotic pressure. F. T. MARTIN AND L. H. SCHULTZ. *J. Phys. Chem.* 35, 638-48 (1931).—The porous disk method of Frazer and Patrick (cf. Townsend, *C. A.* 23, 752) for directly measuring osmotic pressure was modified and applied to dil. solns. of KCl with sufficient success to justify further work on it. Errors in the results were traceable to unequal temps. within the app. and therefore suggest refinements in construction. Distn. was obtained at will in either direction with relation to the disk, which had not been possible in previous work.

H. W. LEAHY

Fractional precipitation. V. The influence of foreign substances in the crystal lattice. OTTO RUFF. *Z. anorg. allgem. Chem.* 195, 60 (1931).—Additional literature reference for a previous article (cf. *C. A.* 24, 2030).

J. B. AUSTIN

Free energy of formation of thallium amalgams. CHARLES EDWIN TEETER, JR. *J. Am. Chem. Soc.* 53, 1100-1 (1931) C. J. WEST

The deposition series of metals in fused salts. GIOVANNI DEVOTO. *Atti III congresso sci. chim. pura applicata* 1930, 327-30—From existing data, the free energy of formation at 700° and at the m. p. was calcd. for hydrides of the alkali metals, alk. earth metals and a few other metals such as *Mn*, *Cu*, *Co* and *Tl*. The potential relative to *H* at 700° was calcd. for a number of metals. J. B. AUSTIN

The diffusion of salt ions thru aluminum. J. CICHOCKI. *Compt. rend.* 191, 841-2 (1930), cf. *C. A.* 23, 5070—An *Al* tube, 2 mm. inside diam., and 4 mm. outside diam. was filled with *NiCl₂*, and drawn to 0.3 mm. outside diam. It was placed concentrically with a tube of *Al* in a larger glass tube, and the system was evacuated. At 600 and 40 v. a thermionic current was obtained, which rose to a max. of about 2.3×10^{-7} amp. in 10 min., followed by a rapid decrease. No neg. emission was obtained. The thermionic current was due to an absorption of the salt by the metal. G. M. P.

The apparent dissociation constants of phenylalanine and of dihydroxyphenylalanine and the apparent free-energy and entropy change of certain amino acids due to ionization. SADAICHI MIYAMOTO AND CARL L. A. SCHMIDT. *J. Biol. Chem.* 90, 163-73 (1931)—The apparent dissociation constants for *D*-phenylalanine and *D*-3,4-dihydroxyphenylalanine were detd. Values for the apparent standard free-energy change for the ionization of those naturally occurring amino acids for which apparent dissociation constants are available were calcd. Values for ΔS , the apparent change in entropy of the system due to ionization, were calcd. for alanine, glycine, arginine, histidine, lysine, aspartic acid and glutamic acid and diiodotyrosine. Calcs. were made which show that no appreciable error results when concn. is used instead of activity in calcs. the primary dissociation constant of aspartic acid and of glutamic acid. F. H. EWERT

Electrolytic water as acid. R. SAXEN. *Chem. News* 142, 49-50 (1931)—A method of prep. *H* is described. *H* is liberated rapidly from pure water at a *Cu* anode with a potential of 20 v. Since the cathode becomes plated with *Cu* a *Cu* cathode is desirable also. *N₂* gas appears at the cathode. White *CuOH* is formed first; this turns through a light blue to a green hydride. Electrolytic reactions cannot be represented accurately by chem. equations. The reaction in question is explained by a rotation of the charged element *H* within *H₂O* groups from the position $OIIIH^+$ to the position ^+HOH under the influence of the charges on the electrodes. The stage of the white hydride is formulated as $HOIIIH^+OIIIH^+$, $OIIIH^+(OIIIH^+)$ and the stage of green hydride as $IIIHOIIIH^+OIIIH^+$, $OIIIH^+(OIIIH^+)$ *Cu*, in which γ represents the pos. and δ the neg. electrode. ALLEN S. SMITH

Buffer solutions from secondary sodium phosphate and citric acid. K. H. SLOTTA AND W. FRANKF. *Ber.* 64B, 452-4 (1931)—Buffer solns. prepd. from these 2 salts have a range of *pH* 2.2 to 7.4. The values given for the various proportions differ considerably from previous reports. V. F. FLARRINGTON

The "drop method" apparatus of Todt for the colorimetric determination of *pH* values. F. M. KOLTHOFF. *Chem. Weekblad* 5, 78-9 (1931)—The several shortcomings of this method are pointed out. Buffer solns. are recommended in order to minimize the chances for making errors. J. C. JIJRRENS

The use of indicators with two color ranges for the photometric determination of the hydrogen-ion concentration. HANS LINSER. *Biochem. Z.* 230, 285-9 (1931)—The use of an indicator with a single color range is preferable to that with 2 color changes in detg. the *pH* by means of the step photometer. S. MORGULIS

The potential of the indo-indochloride electrode. HENRY TERREY AND HERBERT C. BAKER. *J. Chem. Soc.* 1930, 2583-6.—The potential of the indo-indochloride electrode was measured, the cell arrangement being $H_2 | H_2Cl_2 + NaCl | NaCl + Na_2-$

$IrCl_4 + Na_2IrCl_6 | Ir$. The electrodes were of (a) indized *Ir* foil welded to *Pt* *Ir* wire, or (b) indized glass, the *Na₂IrCl₆* concn. being held const. at 0.984×10^{-2} *N* and the *Na₂IrCl₆* concn. being varied from 0.874 to 18.482×10^{-2} *N*. The e. m. f. values vary according to $E = E_0 + RT/F \log \frac{c}{c_0}$, where *c* and *c₀* are the concns. of *Na₂IrCl₆* and *Na₂IrCl₆*, resp. The heat of oxidation of the indo-chloride was calcd. as 5600 cal./mol. It is found that the indo-indochlorides are correctly represented as being of the "ferro-ferri" type. J. BALOZIAN

The reduction potential of thymoquinhydrone. ERNAR BILLMANN AND JYTTE MITS. *Ber.* 64B, 310-4 (1931)—The prepn. of pure thymoquinhydrone is described. Its potential was measured against quinhydrone. The values were not very const., increasing with time and temp. The e. m. f. against *H₂* is at 18° 0.5927 ± 0.0003 v. and at 25°, 0.5806 ± 0.0005 v. V. F. FLARRINGTON

The behavior of the modifications and varieties of silica in a stream of chlorine in the presence of carbon. E. GRUNER AND J. ELÖN. *Z. anorg. allgem. Chem.* 195, 269-87(1931).—The sequence of reactions is shown to be $\text{SiO}_2 + 2\text{Cl}_2 \rightarrow \text{SiCl}_4 + \text{O}_2$, $2\text{C} + \text{O}_2 \rightarrow 2\text{CO}$ and not the more commonly accepted $\text{SiO}_2 + \text{C} \rightarrow \text{Si} + 2\text{CO}$, $\text{Si} + 2\text{Cl}_2 \rightarrow \text{SiCl}_4$. The temps of formation of SiCl_4 from SiO_2 , C and C are 740° from amorphous SiO_2 , 1060° from tridymite, 1060° from cristobalite and 1220° from quartz. For various naturally occurring varieties of SiO_2 , the necessary temp. is 800 – 1100° , depending on the degree of aging. The temp of formation of SiCl_4 can be used in the analysis of SiO_2 to det. the amts. of the various modifications present.

A. P. SACHS.

The influence of the silent discharge on sulfur vapor. ROBERT SCHWARZ AND PAUL ROYFN. *Z. anorg. allgem. Chem.* 196, 1-10(1931).—The behavior of S vapor at a temp. of 460° in a silent elec. discharge was studied by a dilatometric method. A relative contraction was observed. This is explained by postulating the formation of S_2 (thiozone) mols.

L. J. ROSENBAUM.

The existence of hyzone. ROBERT SCHWARZ AND PAUL ROYFN. *Z. anorg. allgem. Chem.* 196, 11-25(1931).—A silent elec. discharge was passed through H_2 at a pressure of 4 cm. Hg and liquid air temp. The decrease in pressure was measured at frequent time intervals. From the amt. of contraction observed and a calcn. of the wall adsorption of H_2 atoms, it is concluded that the effect is due only to the adsorption of H_2 atoms and that there is no evidence of H_2 .

L. J. ROSENBAUM.

The hydrolysis of acetamide. THOMAS W. J. TAYLOR. *J. Chem. Soc.* 1930, 2741-50.—The effect of acids (3, 4, 5 *N* HCl , and 4, 5 *N* HBr and H_2SO_4) and bases (*N* NaOH and KOH) on the velocity of hydrolysis of acetamide (in 0.05 *N* soln.) at 25° is measured, together with the effect of 4 *N* and 5 *N* HCl at 40° and 41° , resp. The salt effects are detd. at 25° with 4 *N* HCl for *N* alkali (Li , Na , K) chloride, 4 *N* HBr for *N* alkali bromide solns., *N* NaOH for 2, 2.5 *N* NaCl , 2, 3 *N* NaBr , and with *N* KOH for 2 *N* KCl , KBr and KI . Curves of the acid normality against the velocity coeff. at 25° and 50° indicate that for HCl the velocity coeff. is max. when the acid concn. is 3 *N*. A similar max. is shown by HBr , but is not present with H_2SO_4 . To explain these max. it is assumed that with acetamide HCl and HBr form complexes ($2\text{CH}_3\text{CO}\text{NH}_2$, HCl and $2\text{CH}_3\text{CO}\text{NH}_2$, HBr , resp.), which do not hydrolyze, while H_2SO_4 does not form such a complex. It is found that with 4 *N* HCl as catalyst, NaCl and KCl accelerate the hydrolysis, while LiCl fails to do so, with HBr , KBr accelerates the reaction, NaBr has no apparent effect, while LiBr retards it. To explain this behavior of salts, the assumption is made that acetamide forms stable complexes with certain salts, which complexes do not hydrolyze. The effect of any salt is the same in bases as in acids.

J. BALOZIAN.

Vapor-phase equilibrium of the simplest cis-trans isomers (dichloroethylene). L. LEBERT AND K. BOLL. *Z. physik. Chem., Abt. A*, 152, 451-2(1931).—The Δp s of cis- and trans- $\text{C}_2\text{H}_2\text{Cl}_2$ are $60.14 \pm 0.04^\circ$ and $47.48 \pm 0.01^\circ$, resp. (purity checked by delec. precision measurements). An exact Δp curve of mixts. of the 2 isomers was detd. Condensation of the vapor in equil. at 300° yielded 63% of cis- $\text{C}_2\text{H}_2\text{Cl}_2$, the thermodynamically more stable form.

FRANK URBAN.

The melting point in the tungsten-rhenium system. KARL BECKER AND KURT MORRIS. *Metallwirtschaft* 9, 1063-6(1930).—The solidus line for the W-Re system was detd. An intermetallic compd., W_3Re , m. 3005° and with 2 eutectics one with 50 atomic % Re and m. 2892° , the other with 67 atomic % Re and m. 2822° , were found.

C. H. LORIG.

The modifications of silver iodide. R. BLOCH AND H. MÖLLER. *Z. physik. Chem., Abt. A*, 152, 245-68(1931).—By roentgenographic and microscopic examn. of AgI prepd. by various methods the phase diagram was constructed. There are 3 modifications, cubic ($< 135^\circ$), hexagonal (135 – 46°) and "hot" cubic ($> 146^\circ$). Methods of prepg. both cold forms are given.

V. F. HARRINGTON.

Phase equilibrium in binary systems with continuous mixed crystals. E. KORDES. *Z. physik. Chem., Abt. A*, 152, 161-96(1931).—A simple empirical equation is given for the phase equil. between crystals and liquid, between crystal modifications within continuous crystals and between the anisotropic liquid state and the isotropic. Previous data are used to prove the general applicability of the equation. The value of this equation in geochemistry (the zone formation in plagioclase, the difference between isopolymorphism and boundary miscibility, etc.), is suggested.

V. F. H.

Phase-rule studies on metallic thiocyanates. 1. The systems $\text{Ba}(\text{CNS})_2$ – NaCNS – H_2O and $\text{Ba}(\text{CNS})_2$ – KCNS – H_2O at 25° . VINCENT J. OCCLESIAW. *J. Chem. Soc.* 1931, 57-60.—The systems $\text{Ba}(\text{CNS})_2$ – $\text{Na}(\text{CNS})$ – H_2O (I) and $\text{Ba}(\text{CNS})_2$ – KCNS – H_2O

(II) are investigated, and the solubilities of NaCNS, KCNS and Ba(CNS)₂ in H₂O are determined at 25°. The ternary diagram of system I shows that no double salt exists at 25°. NaCNS·2H₂O is the stable solid phase at this temp when NaCNS is in eqm with its saturated soln, its transition temp being 30.301° (determined dilatometrically and thermometrically). In system II BaCNS·2H₂O exists over a wide area. It is prepared by allowing a soln containing the correct proportions of the constituent salts to crystallize at 25°, freeing from mother liquor as much as possible by suction, and drying in air in cold weather. At 25° the soln in g/100 g is detd as 58.78 for NaCNS, 70.89 for KCNS and 62.61 for Ba(CNS)₂. J. BALOGH.

Physical constants of the system methane-hydrogen. F. A. TRUTH AND T. H. VERCHOYLE. *Proc Roy Soc (London)* A130, 451-53 (1931), cf C. A. 25, 1422. —The pressure range of Cl₂ isotherms was extended to 215 atm. The m-p curve was detd. The triple point was observed to be -182.41° at a pressure of 8.737 cm. The three phase curve and one isotherm at -182.5° of the system Cl₂-H₂ were detd. F. J. ROSENBAUM.

Kinetics of chlorine bleaches and hypochlorite solutions. JOSEPH J. WEISS. *Z Elektrochem* 37, 20-5 (1931), cf C. A. 24, 5589. —The velocity of spontaneous decomposition of alk. hypochlorite solns by the general reactions $3\text{ClO}^- \rightarrow \text{ClO}_2 + 2\text{Cl}^-$, $2\text{ClO}^- \rightarrow 2\text{Cl}^- + \text{O}_2$, at const. H ion concn is given by $\frac{d[\text{ClO}^-]}{dt} = \left\{ k_1 \frac{f_1}{f_1} + k_2' f_{\text{HOCl}} [\text{H}^+] \right\}$

$[\text{OCl}^-]^2$, $\frac{1}{t} \left\{ \frac{1}{[\text{ClO}^-]} - \frac{1}{[\text{ClO}^-]_0} \right\} = k_1 \frac{f_1}{f_1} + k_2' f_{\text{HOCl}} [\text{H}^+] = k_{\text{lim}}$. Consideration of the F. Fornsier tests (*Z Elektrochem* 23, 137 (1917), cf C. A. 11, 2865) and use of the Debye-Huckel theory gives $k_{\text{lim}} = 1.07 \times 10^{-4} \times 100.704 \sqrt{[\text{OH}^-]} \times \sqrt{[\text{OCl}^-]} + 1.2$

$\times 10^{-4} \frac{f_{\text{HOCl}}}{[\text{OH}^-]}$ at 25°, time in days and concn in mol/l. Consideration of the Kauff-

mann tests (C. A. 19, 561) gives $k_{\text{lim}}^K = (5.1 \times 10^{-4} f_{\text{HOCl}} / [\text{OH}^-])$ in the same units. General kinetics of Cl bleaches, of HOCl or ClO₂ is assumed as the bleaching principle, are set up, and the oxidation mechanism is discussed, with $f(F) = \text{const}$. $\frac{d[\text{ClO}^-]}{dt} = \left\{ f(F) + k_2' \left\{ f_{\text{HOCl}} K'_{\text{HOCl}} [\text{H}^+] [\text{ClO}^-]^2, \frac{1}{t} \left\{ \frac{1}{[\text{ClO}^-]} - \frac{1}{[\text{ClO}^-]_0} \right\} = \left\{ f(F) + k_2' \right\} f_{\text{HOCl}} k'_{\text{HOCl}} [\text{H}^+] = k_{\text{lim}}' \right. \right.$ For homogeneous reactions $f(F) = \text{const}$. $[\text{R}] + k_2' (R \text{ is the reducing agent})$. This is compared with the Kauffmann phenolphthalein tests and confirmed. For heterogeneous reactions a form $f(F) = a_1 [\text{H}^+]^n + a_2$ is discussed, which in general leads to a H ion concn for extreme reaction velocity. There is discussed a reaction velocity and swelling being assumed for the special case of cellulose bleaching or oxidation, where $k_{\text{lim}}' = f_{\text{HOCl}} \{ b_1 \sqrt{[\text{H}^+]} + b_2 [\text{H}^+] \}$. The Clibbens and Ridge tests (cf C. A. 21, 2192) and theory are discussed for the cases $b_1 = 3.92 \times 10^3$, $b_2 = -4.0 \times 10^3$, $[\text{H}^+]^{(n)}_{\text{max}} = 2.5 \times 10^{-7}$, and $b_1 = 10.4 \times 10^3$, $b_2 = -2.2 \times 10^3$, $[\text{H}^+]^{(n)}_{\text{max}} = 0.6 \times 10^{-7}$ as H ion concn of greatest fiber attack. A notable fact for useful application of practical bleaching of cellulose and textiles is that, from the general difference of the n of the fiber and the n of the color to be discharged, a H ion concn or interval can be calculated at which the color to be discharged is attacked rapidly, while the textile or fiber is practically unattacked, which is advantageous as regards bleach loss and textile strength. L. M. SUMMERS.

Influence of the triazo ion in the catalytic action of colloidal platinum on hydrogen peroxide. F. OLIVERI MANDALA. *Gazz chim ital* 60, 878-82 (1930); cf C. A. 24, 1017. —Colloidal Pt was prepared by the method of Bredig. Decomposition of H₂O₂ was followed by N detns with an app used previously for the catalytic decomposition of HN₃ (cf C. A. 11, 1108). The molar concns per l of the H₂O₂, NH₃, and colloidal Pt were 0.0839, 0.012 and 0.00010, resp. The velocity of decomposition of H₂O₂ is reduced greatly by adding small quantities of NH₃, even traces of NH₃ reduce the velocity about 0.5. The velocity const. k , is not a linear function of the concn of NH₃. The presence of HN₃ decreases k about 25%. The temp coeff. K_{440}/K_1 is 1.66 to 2. Conclusion the catalytic decomposition of aq. H₂O₂ follows the logarithmic law $x = A(1 - e^{-kx})$, and the temp coeff. per 10 degrees is approx 2, whether NH₃ is present or not. C. C. DAVIS.

Autoradiation of cyclohexene by oxygen. N. D. ZELINSKII AND P. P. BOARISOV. *Ber* 63B, 2362-5, *J. Russ Phys.-Chem Soc* 62, 2051-4 (1930). —The Lewis Langmuir

theory predicts the addn of O to an ethylene linkage to yield a peroxide without the use of a catalyst. Z and B allowed 20 g of cyclohexene to absorb O. The rate of absorption increased, decreased and became zero after 151 days. 3200 cc of O, corresponding to one atom per molecule, was absorbed. The unchanged cyclohexene was removed by standing in a vacuum desiccator over H_2SO_4 for 3 months, when a tough, resinous, reddish yellow mass remained, which gave the peroxide test with guaiacol resin and ferrous salts, and analyzed for cyclohexene peroxide, $\text{C}_6\text{H}_{10}\text{O}_2$. P. H. RATHMANN.

Oxidation of seemingly autoxidizable leuco bases by molecular oxygen. ALBERT REID. *Ber* 63B, 1929 2(1930).—Thionine was reduced by phenylhydrazine to leucothionine and carefully purified by washing with alc., drying and subliming at 190–220° under 0.05 mm. M. p. 185°. 2.0 mg leucothionine in 2 cc of a molar buffer soln of NH_4OAc and HOAc purified in quartz vessels ($p_n = 4.5$), was shaken in a quartz vessel immersed in a thermostat at 20°, and the rate of absorption of O, noted. From 0.0001 to 0.001 mg of Cu greatly increased the rate. Replacement of the N of the air by CO_2 greatly reduced the rate both with and without the presence of Cu. Leuco methylene blue gave similar results. It is concluded that the oxidation of these leuco bases is not autoxidation but is catalyzed by traces of such metal ions as Cu, occurring as impurity in even the purest of lab. reagents. P. H. RATHMANN.

The catalysis of the reactions between solids. II. The mechanism of the reaction of catalytic stannate formation. SETSURO TAMARU AND NOBORU AYDÖ. *Z anorg allgem. Chem.* 195, 309 20(1931). cf. *C. A.* 24, 774.—If SnO_2 is previously heated to 1150° the stannate formed from the mixt. of SnO_2 and CaO without a catalyst at 900° is always small in amt. and is independent of the method of prepn. of the SnO_2 . The reduction of SnO_2 by H₂ occurs in 2 stages and the further reduction of SnO is repressed by the presence of CaO . A new method of decompose cassiterite for analysis is proposed, based on the formation of stannate by heating with lime in the absence of air. A. P. SACHS.

The decomposition of carbon monoxide in the presence of iron and iron oxides. P. RISCHEITH. *Z. physik. chem. Unterricht* 44, 22–4(1931).—R. describes 2 lecture expts. to study the reaction (a) $\text{CO} \rightarrow \text{C} + \text{CO}_2$ and the equil. (b) $\text{CO}_2 + \text{C} \rightleftharpoons \text{CO}$ in the presence of Fe and its oxides. The gases are passed over heated Fe_2O_3 in a glass tube. Even after reduction of the oxide the content of CO_2 in the escaping gases (a) remains about const. (50%) indicating the decompn. of the CO in contact with Fe. At 400–500° the equil. (b) is about 50% CO_2 . This catalytic reduction of CO may account for the occurrence of the finely divided C so often met with in the upper regions of the blast furnace. E. R. SCHERZ.

The mechanism of iron catalysts in certain oxidations. C. V. SMYTHE. *J. Biol. Chem.* 90, 251–65(1931).—Ferrous iron present as the unionized pyrophosphate or metaphosphate is shown to oxidize rapidly with O_2 of the air regardless of the acidity of the soln., whereas FeSO_4 in similar slightly acid solns. showed no appreciable oxidation. This behavior, as well as the catalytic effect of ferrous and ferric iron in certain org. oxidations, is explained in terms of the electronic structure of the compds. P. H. E.

New methods of organic thermochemistry. M. REBEK. *Archiv. Hem. Farm.* 4, 212 21(221 German)(1930).—A report on preliminary expts. to det. the energy content of org. compds. J. KUČERA.

APPLEYARD, ROLIO. A Tribute to Michael Faraday. London: Constable and Co., Ltd. 204 pp. 7s 6d, net. Reviewed in *Nature* 127, 395(1931).

BARRETT, W. H. The Periodic Law Chart. Comprising 4 separate tables: periodic table and atomic numbers (after Bohr), periodic table and atomic weights (after Mendeleev), periodicity of atomic volumes (after Lothar Meyer), melting points and atomic numbers. London: John Murray. Unmounted, 5s 6d, net; mounted on linen, 9s 6d, net, mounted on linen, varnished, on rollers, 15s 6d, net. 6" × 22".

BUECKER, HANS. Freshman Key and Guide for Examinations in Chemistry. Morgantown, W. Va.: Chancery Row Printing House. 57 pp.

EASON, ALICE B. Flow and Measurement of Air and Gases. 2nd ed., revised. London: Chas. Griffin and Co., Ltd. 254 pp. 20s, net.

POSTER, WILLIAM. Welt und Wunder der Chemie. Munich: Drei Masken Verlag. A-G. 520 pp.

GOARD, A. K. Chemical Composition: The Methods by Which Atomic Weights and Molecular Formulae Have Been Determined. London: Sidgwick & J. 304 pp. 5s, net.

LAKE, CHARLES H., AND UNSELD, GEORGE P. A Brief Course in Physics. New York: D. C. Heath and Co. 464 pp.

SEARS, FREDERICK F. *Essentials of Physics*. 2nd ed. Chicago: Laurel Book Co. 596 pp.

3—SUBATOMIC PHENOMENA AND RADIOCHEMISTRY

W. ALBERT NOYES, JR.

Masses of the electron, the proton and the universe. A. S. EDDINGTON *Proc. Cambridge Phil. Soc.* 27, 15-9 (1931); cf. *C. A.* 24, 1277—L's "Theory of 137" is developed further. The calcn. of the ratio of mass of the proton to the mass of the electron is given as 1819.6. The "packing fraction" is checked. The no. of protons in the universe is given as 7×10^{18} or 14×10^{18} , depending on whether space is taken as elliptical or spherical. WILLIAM E. VAUGHAN

The quantum theory of zero-point temperature. G. BECK, H. BETHE AND W. RIFKIN. *Naturwissenschaften* 19, 59 (1931)—At 0 abs. all internal motion of a crystal lattice has stopped. All degrees of freedom are "frozen" with the exception of 1 per electron for the Bohr orbits of the electrons. Every electron has (Eddington) $1/a$ degrees of freedom if a is the Sommerfeld fine structure const. Likewise $1/a$ degrees for every proton are present. Hence in order to reach abs. 0 ($2/a$)—1 degrees of freedom have to be subtracted from a substance per neutron present. If this is equated to 273°, a value of 137 is found for $1/a$ which is considered to agree with a value obtained in a different way. B. J. C. VAN DER HORVEN

Quantum mechanics of electrons in crystal lattices. R. DE L. KRONIG AND W. G. PENEY *Proc. Roy. Soc. (London)* A130, 499-513 (1931)—Math. The mechanics of electrons in periodic fields of potential are calcd. for an integrable case. The results agree qualitatively with exptl. facts. E. J. ROSENBAUM

Diamagnetism of the free electron. C. G. DARWIN *Proc. Cambridge Phil. Soc.* 27, 86-90 (1931)—Math. D shows that the susceptibility of a free electron gas is $2\mu_B^2/3kT$ per electron, where μ_B is the Bohr magneton. WILLIAM E. VAUGHAN

X- and γ-radiation measurement and the new international unit. J. G. STEPHENS. *J. Cancer Research Comm. Univ. of Sydney* 1, 222-31 (1930)—Radiation is specified clinically by detg. (a) the wave length, (b) the amt. of energy in the radiation and (c) the degree of polarization. The r unit, which measures the total ionic charge under certain conditions, is used to express energy in the radiation in terms of the ionization in air. It is suggested that the ionization in liquids such as hexane or light petroleum would be of advantage by reducing the size of app. and giving larger ionization currents. B. C. A.

Direct measurement of intensity distribution in molecular beams. J. B. TAYLOR. *Z. Physik* 57, 242-8 (1929)—When an alkali metal atom strikes a glowing W wire, it parts with an electron and is reemitted as a pos. ion. By measuring the no. of pos. ions emitted from such a wire in various positions in the path of a mol. beam the no. of alkali metal atoms striking the wire per sec. can be detd., and hence the energy distribution in different parts of the beam deduced, with an accuracy of 1 in 10^3 . B. C. A.

The production of an intense beam of hydrogen positive ions. LOUIS R. MAXWELL. *Rev. Sci. Instruments* 2, 129-40 (1931)—A method was designed for obtaining an intense beam of H_2^+ ions by diminishing the loss of the ions to the walls of the discharge chamber which occurs in the usual method. This is accomplished by placing the discharge chamber in a strong magnetic field in which the direction of the magnetic field is parallel to the direction of motion of the pos. ions. The app. is described. It is capable of producing a pos. H ion beam corresponding to currents ranging up to 3 ma. in a region of pressure of 3.3×10^{-6} mm. of Hg. ALLEN S. SMITH

Photographic detection of the asymmetrical angular distribution of twice-reflected electrons. E. RUPP. *Naturwissenschaften* 19, 109 (1931)—Photographs were made of twice reflected electrons. An electron beam passed through a 0.3-mm. hole and was reflected at 45° on a Au surface. The reflected ray (also at 45°) passed through a 0.1-mm. hole (50 mm. from the point of reflection) with a thin Au foil behind it. The scattered ray is photographed on a plate 450 mm. from the foil. For electrons of 220 kv., diffraction rings were obtained with distinct asymmetry, the rings being weakest on the side away from the original direction of incidence of the electron beam. The asymmetry is not due to the structure of the Au. An Al reflector and foil do not give the same result. B. J. C. VAN DER HORVEN

Angular scattering of electrons in gases. F. L. ARNOT *Proc. Cambridge Phil. Soc.* 27, 73-6 (1931); cf. *C. A.* 24, 1572—A discussion is given of the error introduced

into exp's on the angular scattering of electrons in gases by a potential at right angles to the electron beam. Calen shows that in the previously reported paper the error in all cases was less than $0^{\circ}17'$, which is less than the exp'd error in the det'n of the angles.

WILLIAM E. VAUGHAN

Optical determination of the sphere of action of atoms for electrons. L. S. ORNSTEIN AND A. M. VAN DOORNHUIZEN. *Proc. Acad. Sci. Amsterdam* 33, 683-9 (1930), cf. *C. A.* 24, 1793, 2069.—The optical method previously described for the det'n of the sphere of action of He atoms is limited to electronic energies above the excitation potential. It is shown that it is possible to extend the measurements below that value with a mixt. of He and Hg. Further it is possible to simplify the method so that intensity measurement is not necessary and the density det'n. is sufficient for obtaining the value of the sphere of action.

MARIE FARNSWORTH

The cathode radiation of the sun. H. RUDOLPH. *Naturwissenschaften* 19, 66 (1911).—I coin a theory on the daily variation of the earth's magnetism (*Gerl. Beitr. z. Geophysik* 27, 378 (1930)) it is concluded that the N "dust shell" around the earth is negatively charged. This necessitates a strongly neg. corpuscular radiation from the sun and a potential of some 10^{11} v. for this shell. Several conclusions and results are discussed.

B. J. C. VAN DER HOEVEN

Direct photography of ionization in insulating substances. A. GEMANT. *Naturwissenschaften* 19, 109 (1931).—Light-sensitive films in an elec. field are blackened on reaching certain limits of potential. This effect is attributed to ionization, it occurs in gases as well as in liquids and solids. By using photographic paper (semiconductor) as an extension of the plate electrode the structure of insulating substances can be examined when present between the electrodes. Blackening below the breakdown potential indicates porous structure of the material, ionization and consequent luminescence by either d. e. or a. c. can thus be studied directly. It was found that for a no. of substances only a. c. gives the glow to which the blackening is due.

B. J. C. VAN DER HOEVEN

Photoelectric effect with lead and mercury at low temperatures. J. C. McLENNAN, R. G. HUNTER AND J. H. McLEOD. *Trans. Roy. Soc. Can.* [3], 24, Sect. 3, 3-23 (1930).—Measurements of the photoelec. current produced by films of Pb and of Hg when irradiated with ultra-violet light below 3000 Å. U. were made down to temps. below the transition point for supercond. The object was to det. whether the emission of electrons from a metal when suitably irradiated undergoes any abrupt modification if the temp. of the metal irradiated is progressively lowered from a point a few degrees above its supercond. transition point to one a few degrees below it. The photoelec. current from Pb and Hg in highly evacuated cells was measured at temps. down to that of liquid H₂. At the temp. of liquid air Pb showed a decrease in the current of about 2% from that of the sat'n. value of the current at room temp. and a decrease of about 16% at the temp. of liquid H₂. Hg at liquid air temps. showed large variation with the thickness of the Hg film. Inability to reproduce films of exactly the same thickness made it impossible to draw any conclusions concerning the change in the photoelec. current from Hg at the temps. of liquid air and liquid H₂. The photoelec. current from lead films at the temp. of liquid He and in the superconducting state was slightly less than it was at a few degrees above its transition temp., 7.2°K. No abrupt change took place in the photoelec. emission of electrons from Pb as it was cooled down through its characteristic transition temp.

J. W. SHIPLEY

The action of adsorbed gas films on the photoelectric effect of salts (remark on "The action of Geiger counting chambers"). J. KLAMMECKE. *Naturwissenschaften* 19, 87 (1931).—Water vapor promotes electron emission in a manner similar to that found in Geiger counters by Bosc and Klumb (*C. A.* 25, 1434) in an investigation of the photoelectric effect of salts. Salts heated *in vacuo* show no photoelectric effect (Werner, *C. A.* 23, 1295). Salts sublimed in air give a photoelectric current, which increases 10 fold after a short treatment with water vapor. They do not exhibit the same activity as do the same salts with occluded water prep'd from solus. Results on gases other than H₂O are inconclusive.

B. J. C. VAN DER HOEVEN

Cathode sputtering at very low gas pressures. A. GÜNTHERSCHULZE AND K. MAYR. *Z. Physik* 62, 607-18 (1930).—The sputtering of Ag and Cu cathodes in He, Ne, Ar, H₂ and N has been investigated at pressures so low (generally less than 0.01 mm., in the app. described) that the sputtering is not dependent on the gas pressure. The sputtering electrode, changes in the wt. of which were noted by observing the motion of a spiral spring by which it was supported, was placed near a low-voltage arc consuming 3-5 amp. at 50 v., it was maintained at a p. d. of 300-1200 v. neg. with respect to the anode of the arc. The amt. of sputtered metal increases with the p. d. applied

to the cathode, and the amt per unit ion current reaching the cathode increases in the order $\text{H} < \text{Ne} < \text{N} < \text{A}$, at any given cathode potential. No sputtering was observed in H up to the highest p d applied, 800 v. The efficiency of the sputtering calcd as the ratio of heat of vaporization of the metal sputtered to the energy delivered by the ion current ranges from 0.3% in H to 1.3% in A . Data have also been compiled relating to the amt of gas adsorbed per g of metal sputtered under various conditions of gas pressure and sputtering electrode potential, the values obtained are in the neighborhood of 0.4 mg in A , 1.0 mg in Ne , 0.7 mg in H and 1-6 mg in N . B. C. A.

Observation of a particulate emission from the cooled metallic cathode of an electric arc. GEORGES LAUDET. *Compt rend* 192, 202-3(1931) M. McMAHON

The dielectric constant and the conductivity of ionized gases. TH. V. JONESCU AND C. MIHUL. *Compt rend* 192, 343-5(1931).—The values of the cond and the dielec const of electronic gas as detd previously (C. A. 25, 1440) were about 30 times greater than the calcd values. The present study indicates that this was due to the assocn of electrons with mols to a degree depending on the elec. conditions present in the tube used for the measurements. The results suggest that the long transmission range of short elec. waves of between 10 and 40 m wave length might be due to a similar assocn of electrons with mols in the upper atm. The reflecting power of such an ionized layer would be great for wave lengths between 10 and 40 m. R. H. LOMBARD

Method of measuring the number of ions in the free atmosphere. YO. IITWARA. *Phyisk Z* 32, 97-106(1931).—The usual type of ion-counting tube does not measure the total no. of ions in the air as some are deflected away from the entrance by the elec field. Measurements were made of the contour surfaces of equal potential around the end of the counting tube by means of a flame probe. I. discusses and proposes corrections to apply to the data of Hess (*Sitzb Akad Wiss Wien, Abt IIa*, 138, 109 (1929)). Reply V. F. Hess. *Ibid* 106.—H disagrees with I. on several points. T. H. CHILTON

Experiments on the entanglement of atoms in a magnetocathodic or cathodic stream. E. HENRIOT, O. GOCHÉ AND (Mlle) F. DONDY-HÉNAILT. *J phys radium* [7], 2, 1-11(1931).—Atoms of various substances (W, C, Pt, Na and S) may be entangled in a magnetocathodic or cathodic stream, probably as ions. Once entangled in the electron stream they appear to form an integral part of it in that they are deformed by an elec or magnetic field in the same way and yield deposits of the particular element used of the same shape as the electronic stream alone. They resist various attempts to sep them from the electronic stream. However, they may be disengaged from the electronic stream at points of strong curvature of the latter and are obtained as pos ions. This is brought about by applying an opposite and larger voltage between the plates than that of the electron flow. The electron stream is then bent back to the plate through which it came after reaching a point of great electron density where the curvature of the stream reaches a max, whereas the atoms continue and are deposited on the opposite plate. No satisfactory explanation of the attraction of the electron stream for the atoms is advanced but is thought possible that it is of an electrostatic character. OGDEN E. SHEPPARD

The anomalous scattering of α -particles by light nuclei. EUGEN GUTH AND THEODOR SEHL. *Z Physik* 66, 577-80(1930).—The theory and formulas for the exact calcn of anomalous scattering of α particles in a Gamov Gurney-Condon potential field are described. M. McMAHON

A new method for the radioactive investigation of pulverized substances. W. SEBESTA. *Z Physik* 66, 598-612(1930).—Because of its great sensitivity the α ray method is suited to quant measurement of radioactivity. In the method described a small quantity of pulverized normal material was mixed with the pulverized material to be investigated and the resulting change of gas cond. produced in an α ray counter was measured. Ionization took place in a semi spherical condenser with circular opening before which the brass dish holding the pulverized material was placed, a quadrant electrometer was used as measuring instrument. The standard material used was Ra Ba sulfate. M. McMAHON

An investigation for accomplishing the decomposition of lead atoms. III. A. SMITS AND H. S. VENING MEINESZ. *Proc Acad Sci Amsterdam* 33, 737-48(1930); cf C. A. 19, 3210, 21, 855. MARIE FARNSWORTH

Artificial disintegration by α -particles. J. CHADWICK, J. E. R. CONSTABLE AND E. C. POLLARD. *Proc Roy Soc (London)* A130, 463-89(1931).—Atoms of B, N, Al, P, Na and P were bombarded by α particles from Po. The protons emitted were recorded by an elec method with vacuum tube amplification. It is assumed that the

protons and α particles are in definite energy levels within the nucleus. Mass de have been calcd and checked against Aston's data (*C. A.* 21, 3513). E. J.

Photoelectric absorption of γ -rays. L. H. GRAY *Proc. Cambridge Phil. Soc.* 103-12(1931).—Consideration of the exptl data has lead G to the following eqn for the photoelec. absorptn coeff per electron (τ) of x rays and γ rays: $\log_{10} \tau = 3 + 1.0 \log_{10} \lambda + 0.150 (\log_{10} \lambda)^2$. The curve is given. WILLIAM F. VAUGHAN

Capture of electrons by α -particles. H. C. WEBSTER *Proc. Cambridge Soc.* 27, 116-30(1931), cf *C. A.* 24, 5617.—The study of the capture by α par from Po of electrons emitted from a heated oxide covered filament was attempt. The ions were recorded with a modified Geiger counter. The exptl set up is desc in detail. Wholly neg results were obtained. The disagreement with the wor Davis and Barnes (*C. A.* 24, 1029) is discussed, no reconciliation is made. W. B.

Study of the velocity of disintegration of polonium in various places. L. B. VALENSKII *J. phys. radium* [7] 2, 12-9(1931), cf *C. A.* 24, 3143, 24, 3427.—7 series of expts on the detn of the half period were made in the southern part of S. R. (42 places) in the region of Leningrad (18 places) and at Dietkovo Selo (3 pls). The results vary from 135.5 to 141.1 days. The time seems to depend on the where the detn was made. L. D. ROBER

The nature of the absorbable radiation accompanying the rays from polon. IRENE CURIE AND ERDREIC JOLIOT *J. phys. radium* [7], 2, 20-8(1931).—A detailed account is given of the work described in *C. A.* 24, 1792. O. E.

Theory of the scattering of short x-rays by molecular hydrogen. H. S. W. MA *Proc. Cambridge Phil. Soc.* 27, 77-85(1931), cf *C. A.* 25, 849.—Theoretical, on basis of wave mechanics. M develops a theory which leads to the conclusion tht scattering of short x rays by mol H is similar to that for at H, although the varies with the angle of scattering. There is no agreement with the one set of values available (cf Barrett, *C. A.* 22, 1275, 3581). WILLIAM F. VAUGHAN

Moseley diagram of the ionization voltages of the light atoms and ions. BRAUNNEK *Z. Physik* 63, 20-9(1930).—The sq roots of the ionization voltage the elements and ions of the 1st 3 short periods (to K) are plotted against the at of the element as abscissa, and it is shown that each isoelectronic series (e , Li , B^{++} , C^{+++} , ...) gives points lying on a straight line, the gradient of these lines const in each period. The only point deviating appreciably is that correspond with Cl^+ , and for this ion the ionization voltage of 22.2 v. is predicted instead of provisionally accepted value of 1832 v. It is shown that this means that a relation exists between the screening nos and the atomic no Z in an isoelect series, although in the 1st period (H and He) the screening no is independent of nuclear charge. It is shown that s decreases with increase in Z in an isoelect series, contrasted with the x-ray terms, in which s increases with increase. A slight concavity in the lines of the Moseley diagram points to a decrease with Z which is not exactly linear, the concavity being greatest in the 2nd period not observable in the 1st period. On this basis values for the ionization potent the ions Li^+ , Be^{++} , B^{+++} , C^{+++} are extrapolated from the known value of s for He, com linearity being assumed. These values agree to within 1% with the values calc by Hylleraas (*C. A.* 24, 1235) by means of wave mechanics, which agree very ex with the unpublished measurements of Edlén and Eriksen (cf *C. A.* 24, 1575). The tion of Hylleraas is shown to predict a slight decrease of s with increase in the mentioned series. B. C.

Satellites of the $K\beta_1$ line of elements from iron to zinc. SUKICHI KAW *Mem. Coll. Sci. Kyoto Imp. Univ.*, Ser. A, 13, 383-7(1930).—In the course of exp det the relative position of the absorption edge and the $K\beta_1$ line for the elements Fe to Zn several new satellites have been found on the short side of $K\beta_1$ and $K\beta_2$, seem to correspond to β'' and β''' observed by Druyvesteyn for elements of at no than that of Fe. Wave length tables are presented. The application of theories of Druyvesteyn and Beuthe as to the origin of these lines is discussed. A definite conclusion is reached. C. J. HUMPHRI

X-rays in the service of chemistry and industry in 1931. GEORGE L. C. *J. Chem. Education* 8, 625-39(1931). E.

Experimental study of the absorption formulae of x-rays. H. MATAXICHI I AND SUKICHI KAWATA *Mem. Coll. Sci. Kyoto Imp. Univ.*, Ser. A, 13, 375-81(1931), cf *ibid* 10, 311(1927).—The relation of the true μ -atomic absorptn coeff. $\tau/\rho A$, to at no, Z and to the wave length has been studied for a no. of elements by the method described in the previous paper (*C. A.* 22, 3090). From a study of the data thus obtained it seems right to conclude that the fluorescent absorption of x-rays as a fun

of λ and Z does not obey such a simple law as given by $r/p A = c\lambda^2 Z^2$, in which c , p and g remain const., but that it must be connected by a more complex relation. Also, if it is to be expressed by the above formula, the constancy of p and g should hold only in a restricted region, and in the general case, they should depend somewhat on λ and Z .

C. J. HUMPHREYS

A new x-ray spectrograph for wave-length determinations in air. S. ZEIDENFELD
Rev. Sci. Instruments 2, 153-63 (1931)

E. H.

The accuracy of interference measurements in the molecule with Röntgen and cathode rays. L. BEWILGUA. *Physik. Z.* 32, 114-7 (1931).—From a consideration of the theoretical 'form factor' curves for Röntgen rays and for electrons, it is shown that, in general, the former must give more exact values for the distance between atoms in a single mol.

T. H. CHILTON

Interferometer measurements in the infra-red region of neon and iron. J. C. McLENNAN AND FLORENCE M. QUINLAN. *Trans. Roy. Soc. Can.* [3], 24, Sect. 3, 47-52 (1930).—An examn. of the infra red region of the spectra of Xe and Fe was made to det. whether suitable lines might be found as secondary standards of wave length. Difficulty was found in reproducing the Xe spectrum because of a change of the pressure of the gas as the discharge tube became older. This makes the gas unsuitable as a source of accurate standard wave lengths. A table contg. the wave lengths of the arc spectrum of Fe from 26077.991 to 8688.637 A. U. is given.

J. W. SHIPLEY

Interpretation possibilities of the mercury hyperfine structures. H. SCHÜLER. *Naturwissenschaften* 18, 895 (1930).—The Hg lines, in general, appear to consist of one strong component and several weaker components so grouped that their center of gravity coincides roughly with that of the main one. The ratio of the intensity sum of the weaker components to the total intensity has been measured for Hg arc lines 4047, 4078, 4916, 5461, 5769, 5791 A. U. and for Hg spark lines 6150 and 7044 A. U. With 2 exceptions the ratio is between 0.25 and 0.32. According to Aston the odd numbered isotopes constitute about 30% of the total so that the large hyperfine structure splitting of Hg lines may be ascribed principally to the odd numbered isotopes, Hg¹⁹⁹, Hg²⁰¹.

W. P. MEGGERS

Photoelectric intensity measurements in the mercury spectrum. H. L. S. ORNSTEIN AND J. P. CUSTERS. *Proc. Acad. Sci. Amsterdam* 33, 809-13 (1930).—In the first paper (cf. *C. A.* 24, 4458) it was stated that a characteristic relation of the line intensity I as a function of the current i exists at const. pressure, p , provided that the current is small and the pressure lies within a certain range. Further measurements show that I/i increases with decreasing i at all pressures, provided the current becomes sufficiently small.

W. I. MEGGERS

Intensity measurements for the multiplet s^4G-e^4P . W. A. M. DEKKERS AND A. KRUTHOF. *Z. Physik* 66, 491-3 (1930).—Intensity measurements of 6 lines in the Ni spectrum forming the multiplet s^4G-e^4P (4592.532 to 4900.97 A. U.) were made under a variety of conditions. They confirm the deviations from the sum rule reported by Ornstein and Bouma (*C. A.* 24, 5620).

W. I. MEGGERS

Note on a phenomenon connected with the aurora. A. C. BURTON. *Can. J. Research* 4, 52-3 (1931).—A description with photographs is given of a brilliant white beam of light stretching across the sky from east to west on the night of Aug. 21, 1930. Microscopic examn. of the photograph showed a fine structure not observed visually.

J. W. SHIPLEY

Excitation functions in the neon spectrum. W. HANLE. *Z. Physik* 65, 512-6 (1930).—The intensities of various Ne lines as a function of the applied potentials which excite them are measured in the yellow and red, the potentials ranging from 20 to 100 v. Lines involving the same initial term are found to have the same excitation function. The marked dependence of intensity distribution in the spectrum upon gas pressure is ascribed to an effect of collisions.

W. I. MEGGERS

Remark on our work "Intensity measurements in the copper arc." L. S. ORNSTEIN AND D. VERMEULEN. *Z. Physik* 66, 490 (1930).—In the work referred to (*C. A.* 25, 249) the intensity ratio of the Cu doublet 1^1S-2^1P was reported to deviate from the sum rule. Further measurements made with C electrodes contg. a trace of Cu so that the lines in question could be photographed with a long exposure led to the true value of the intensity ratio, viz., 1/2.

W. I. MEGGERS

Absorption spectra of dissolved mercury. H. REICHARDT AND K. F. BONHOEFFER. *Z. Elektrochem.* 36, 753 (1930).—The soly. of Hg in water at temps. below 130° is too small for the absorption spectrum of the metal to be detected but at about 140° 2 absorption bands of about 20-30 A. U. width, due to the Hg atom, occur at 2600 and 2520 A. U. They are not derived by broadening and displacement of the normal lines

at 2650 and 2537 Å U., since in the solvents MeOH and $C_{12}H_{10}$, both of which contain dipoles, the bands occur at 2575 and 2530 Å U. and 2575 and 2515 Å U., resp.; they may result from division of the line at 2537 Å U. by the action of the elec. fields surrounding the solvent molecules. H. C. A.

Interpretation of band spectra I, IIa, IIb. ROBERT S. MULLIKEN *Rev. Modern Physics* 2, 693-8(1930), cf *C. A.* 24, 2049 W. F. MINGOS

Interpretation of band spectra, IIc. Empirical band types. ROBERT S. MULLIKEN *Rev. Modern Physics* 3, 89-155(1931) E. H.

Some bands of the carbon molecule. G. H. DREKE and W. LOCKYER-HOLTGRVEN *Z. Physik* 62, 787-91(1930)—Details are given for new bands resembling the Swan bands of C_2 (cf *C. A.* 24, 2050) W. F. MINGOS

Radiation and characteristics of molecules. R. D. KLENNAN *Z. anorg. allgem. Chem.* 195, 161-72(1931)—A general discussion WILLIAM E. VAUGHAN

Structure and spectra of the molecules of hydrogen and helium. W. WENZEL *Z. Elektrochem.* 36, 509-603(1930), cf *C. A.* 24, 2160—A description is given of the manner whereby the band spectra of the H_2 and He_2 molecules may be developed by consideration of the arrangement of the electrons and protons, and the phys. significance and method of detg. the quantum nos. of the individual electrons are indicated. The origins of the ortho- and para- H_2 spectra by sym. and asymmetrical rotations are described. H. C. A.

Energy levels of molecular oxygen. J. C. MCLINNAN, H. D. SMITH and J. O. WILLIAMS *Trans. Roy. Soc. Can.* [3], 24, Sect. 3, 65-70(1930)—The absorption spectrum of O_2 was studied in the gaseous, liquid and solid phases to det. the origin and relation of the bands observed in its spectrum. Particular attention was paid to the bands in the spectrum of liquid O_2 . Wave length tables and mol. configuration diagrams are given. J. W. SIMPSON

The band spectrum of silver hydride. FRANK BAYLTON *Nature* 127, 11(1931)—An improved source has greatly extended the range of the known spectrum. This source is an elec. arc operating in H_2 at reduced pressure and having a pos. electrode made of a Ag-Al alloy. Greatly increased intensity has been attained as well as the elimination of the N_2 spectrum. The rotational structures of 14 bands belonging to the $^2\Sigma \rightarrow ^2\Sigma$ system have been analyzed and arranged in a vibrational scheme. The vibrational levels of the lower electronic state appear to be represented by the formula $\bar{\nu}(v) = 1721.5v^2 - 23.5v - 0.0091v^4$. The excited electronic state shows irregularities of spacing in its rotational and vibrational levels which may be due to a perturbing electronic level. An approx. calcn. of the dissociation energy in both states gives $D^0 \sim 19,000 \text{ cm}^{-1}$, $D^* \sim 19,000 \text{ cm}^{-1}$. The isotope effect is clearly shown with high dispersion. C. J. HUMPHRIES

Fluorescence of mercury vapor under atomic and molecular absorption. RAYNIGHT *Nature* 127, 10(1931)—The excitation of Hg vapor by radiation near the resonance line is due to both at. and mol. absorption. The at. absorption gives a type of fluorescence known as the core effect, since the radiation capable of this kind of absorption is limited to about 0.05 Å U. on either side of the resonance line. The intensity of this effect falls off very rapidly with penetration. The weaker fluorescence due to outlying radiation is called the wing effect, and is caused by mol. absorption. It is extinguished much less rapidly as the beam traverses the vessel. The reality of both effects has been confirmed by showing that a suitable adm. of H_2 suppresses the core effect. C. J. HUMPHRIES

The vibrational levels of the iodine chloride molecule. W. I. CURTIS and O. DARYNISHED *Trans. Faraday Soc.* 27, 77-87(1931), cf *C. A.* 24, 1020—Const. pressure measurements of the spectrum were made at 17° and 237° to compare the intensities. All of the vibrational progressions increased appreciably, invalidating the method for the detn. of vibrational levels. Some evidence indicated that the 2nd progression might remain const. and be identified with $v' = 1$. ARTHUR TREISCHER

Raman effects with liquid and gaseous nitrous oxide. J. C. MCLINNAN, H. D. SMITH and J. O. WILLIAMS *Trans. Roy. Soc. Can.* [3], 24, Sect. 3, 107-205(1930)—The Raman effect was studied for liquid N_2O as an example of a triat. mol. and the results were compared with those obtained by others. Tables and a discussion of the results are included. J. W. SIMPSON

Infra-red spectroscopy. M. CERNY *Z. Elektrochem.* 36, 616-8(1930)—A historical survey of the development of modern methods of infra red spectroscopy, with some exp't. details. H. C. A.

Raman effect and its significance for the spectroscopic study of molecular structure. A. SIEGAL *Z. Elektrochem.* 36, 618-31(1930)—A survey of recent work. H. C. A.

Raman effect and chemical bonds in certain organic liquids. LESLIE E. HOWLETT. *Can J Research* 4, 79-91(1931)—Spectrograms were taken of ethylene glycol and 5 of its derivs., 4 nitriles and benzyl alc. on a specially constructed spectrograph designed to study the Raman effect. By assuming the simple harmonic oscillator theory an attempt was made to assoc. certain frequencies with definite bonds and structures. The theory was applied through an expression derived for the frequencies characteristic of the elastically bound masses vibrating in a straight line. The value of the stretching force of all single bonds is assumed to be the same. Double and triple bonds are taken as having a stretching force 2 and 3 times, resp., that of the single bond. Frequencies in the neighborhood of 300 mm^{-1} are attributed to longitudinal vibrations between C and H. Lines near 160 mm^{-1} indicate a double bonded C atom in the chain or ring. Lines near 225 mm^{-1} are due to the $\text{C}\equiv\text{N}$ triple bond. Lines between 50 mm^{-1} and 150 mm^{-1} are due to the longitudinal vibrating chain. Frequencies less than 50 mm^{-1} are probably due to transverse vibrations of the chain, while some lines near 140 mm^{-1} may be due to transverse vibrations of H. In the ethylene glycol compds. lines near 80 mm^{-1} are assigned to the presence of O at the end of the chain. Lines near 72 mm^{-1} and 65 mm^{-1} are attributed to the longitudinal vibrations of Cl and lines near 42 mm^{-1} and 30 mm^{-1} are attributed to transverse vibrations of Cl.

J. W. SIMPLEY

The influence of absorption of light on the rate of photochemical reactions. A. K. BHATTACHARYA AND N. R. DHAR. *Z anorg allgem Chem* 196, 26-32(1931), cf. C. A. 24, 553—The relationship between absorption of light and reaction rate was obtained for the following photochem. reactions: $\text{K}_2\text{C}_2\text{O}_8$ with I, FeSO_4 with I and citric acid with H_2CrO_4 . The exponent in the equation relating absorption of radiation to reaction rate can vary from a proper fraction to 2, it is independent of the relationship between the light and dark reaction. Tabulated data are given. E. J. ROSENBAUM

Photochemical dissociation of triatomic molecules. II. Potassium cyanide. DONALD S. VILLARS. *J Am Chem Soc* 53, 405-11(1931), cf. C. A. 24, 1034—KCN vapor at 875° was found to absorb in 2 regions, the long wave length limits of which are approx. 2175 (5.7 v) and 2900 (4.3 v) Å. The former region is interpreted as corresponding to the dissociation into a normal atom and an excited one (K excitation 1.6 v, CN excitation, 1.8 v), the latter corresponds to dissociation into 2 normal atoms or radicals. This agrees with theoretical considerations. WILLIAM E. VAUGHAN

Photochemical studies. XII. Photochemical reaction between nitric oxide and mercury vapor. W. ALBERT NOYES, JR. *J Am Chem Soc* 53, 514-26(1931)—The reaction between NO and Hg in the full radiation of a quartz Hg lamp immersed in water has been studied from pressure decrease in the system. The reaction seems to lead to the formation of NO_2 , which may react with Hg to give N_2 . Both quartz and Correx glass cells were used, the pressure of the NO was initially approx. 0.1 mm. The rate in all cases was computable on the basis of a monomol. reaction rate equation. Initial pressures were varied over the range 0.1125 to 0.00338 mm. and the monomol. rate still applied. The value of k_1 decreases with high pressures of the order of 24 mm. Removal of Hg vapor practically stopped the reaction. Some studies on the quenching of the fluorescence of Hg by NO were made, 72% quenching at 2 mm. and 85% at 6 mm. A detailed theoretical discussion is given, various mechanisms are discussed. It is thought that a NO mol. in a high vibration state of the normal electron level is produced by collisions of the 2nd kind.

WILLIAM E. VAUGHAN

The photosensitized decomposition of nitrogen trichloride and the induction period of the hydrogen-chlorine reaction. J. G. A. GRIFFITHS AND R. G. W. NORRISH. *Nature* 127, 14(1931)—The inhibiting effect of the addition of small measured quantities of NH_3 on the induction period of the photochem. HCl reaction is due to the formation of NCl_3 , which undergoes photosensitized decomposition in the presence of excess Cl. Quantum-efficiency measurements indicate that the decomposition of NCl_3 proceeds by way of reaction chains of short length, the quantum efficiency falling to a lower limiting value of about 2 as the Cl pressure is increased. Gases such as He, Ar, N and O, when added to mixts. of NCl_3 and Cl_2 have sp. retarding effects on the decomposition; and, as the pressures are increased, depress the quantum efficiency toward the above limiting value.

C. J. HUMPHREYS

The absorption of aqueous tartaric acid solutions. RENE LUCAS AND MARCEL SCHWOB. *Compt rend* 192, 225-7(1931)—The relative results of ultra violet absorbing powers of aq. tartaric acid solns. were detd. with the help of photoelec. cells, the photoelec. cell arrangement of Halban and Siedentopf (C. A. 16, 2078) being used. Contrary to previous observations, Beer's law was not verified. The variations are

particularly marked for 2536 Å U, the sp absorption varying more than 200%. Dil. solns having greater absorbing power than concd solns M. McMAHON

Absorption spectrum of lignin derivatives in the ultra-violet. ERIK HÄGGLUND AND F. W. KLINGSTEDT *Z. physik. Chem., Abt. A*, 152, 293-312 (1931); cf. *C. A.* 24, 1578—The absorption spectra of the alkyl lignins, the ligninsulfonic acids and the alkali lignins have been measured by the method of Victor Henri by means of the small Hilger quartz spectrograph. The general similarity of the spectra of the different substances points to the fact that they have a common residue, there is, however, a distinct difference between the spectra of derivs of lignin from conifers and from deciduous trees. Selective absorption is relatively strong which indicates the existence of at least 1 aromatic ring in the fundamental substance of lignin W. WEST

Potassium chromate and potassium dichromate as light filters and the constitution of chromic acid from absorption measurements. W. V. BHAGWAT AND N. R. DHAR. *J. Indian Chem. Soc.* 7, 913-21 (1930)—Data of extinction coeffs shows that wave lengths shorter than 5150, 4950, 4930, 4600, 4300 and 4000 Å U are completely absorbed by 1 cm of 5.5 (satd), 2, 1, 0.1, 0.01, and 0.001 N K_2CrO_4 solns, resp. wave lengths shorter than 5600, 5200, 4600 and 4000 Å U are completely absorbed by 1 cm of 1.17 (satd), 0.1, 0.01 and 0.001 N $K_2Cr_2O_7$ solns, resp. The % transmission in the range 5600-6700 Å U is greater with satd K_2CrO_4 than with satd $K_2Cr_2O_7$ soln. In 0.001 N solns, K_2CrO_4 transmits light of wave lengths 5400-6700 Å U better than $K_2Cr_2O_7$. Neither K_2CrO_4 nor $K_2Cr_2O_7$ solns obey Beer's law. In dil solns the absorption-concn curves for K_2CrO_4 and $K_2Cr_2O_7$ are strikingly similar. Although concd solns of dichromate contain $Cr_2O_7^{--}$ and $HCr_2O_7^-$ ions, the dil solns contain mainly CrO_4^{--} and $HCrO_4^-$ ions O. CAR T. QUIMBY

The law of paramagnetic rotation of xenotime and its experimental verification. JEAN BECQUEREL, W. J. DE HAAS AND H. A. KRAMERS. *Compt. rend.* 191, 839-41 (1930), cf. *C. A.* 24, 2047. GERALD M. PETTY

Quantum theory of chemical kinetics. Unimolecular reactions. S. ROBINSON AND L. ROSENKEVICH *Z. physik. Chem., Abt. B*, 10, 47-85 (1930)—A unimol decomn or rearrangement can be viewed either as originating in an internal force field within the complicated mol., which causes the expulsion of a part of the mol. in a manner analogous to the mechanism of γ -ray disintegration as pictured by Graw and by Gurney and Condon, or as a rearrangement of its parts whereby new chem linkages are formed and simultaneously a part of the old linkages is disrupted with the consequent emission of part of the original mol. This latter process has its analog in the Auger effect, as e. g., in the excitation and simultaneous spontaneous ionization of a He atom. Both schemes lead to the correct exponential form of relation between velocity const. and temp.; but it is argued that there is no expl. evidence for the kind of interaction between the mol. and the emitted particle necessary in the first scheme. The Auger process gives $k = e^{-A/RT} e^{\beta(Q-A)/\sqrt{A}}$, where A is the energy of activation, Q the heat of the reaction, and β a factor depending only slightly on the nature of the mol. This requires an exponential relation between the consts A and B of the Arrhenius equation $k = Be^{A/RT}$, which is shown to be experimentally satisfied. The order of a reaction is detd. by the relative values of Q and A ; the conclusions are applied to known cases and used in predictions concerning reactions for which data are still wanting W. WEST

Chemical combination and the line emission of solid bodies. R. TOMASCHKE *Z. Elektrochem.* 36, 737-43 (1930)—A survey of work on the spectrum of Sm in phosphors. Cf. *C. A.* 18, 3546; 19, 935; 22, 915 B. C. A.

The ionization of air during the oxidation of phosphorus. J. TAU SZ AND H. GÖRLACHER *Physik. Z.* 32, 91-7 (1931)—Ionization measurements were carried out incidentally to work on the oxidation of P in air and O_2 , both pure and in the presence of poisons for the reaction (*C. A.* 24, 4476). When P was oxidized by air, the ionization increased with air velocity, temp., and P surface. Pure O_2 , when wet, showed no increase in ionization over air, but showed a progressive increase during drying. Strongly poisoning substances such as isoprene and cyclohexane mixed in air caused considerable decreases in ionization depending on the concn. T. H. CHILTON

Physical and chemical action of ultra-violet rays on sublimed S (FONZES-DIAZOV)
15. Behavior of water with change of temperature and with addition of electrolytes as studied by the Raman effect (RAO) 2. Determining Hg vapor in air (Ger. pat. 517,480) 7.

- KAYSER, H AND KONEN, H: *Handbuch der Spectroscopie*. Band VII, Lfg. 2. Leipzig S Hirzel Pp 499-750 M 28
- LODGE, OLIVER *Atoms and Rays: Modern Views on Atomic Structure and Radiation* London Benn 222 pp 3s 6d net
- PETTERSSON, HANS *Teneur en radium des dépôts de mer profonde*. Monaco: Imprimerie de Monaco 50 pp

4-ELECTROCHEMISTRY

COLIN O PINK

Recent progress in the use of the electric furnace in the steel industry. ALBERT LEVASSEUR *Arts & métiers* 1931, 54-60, cf C A. 24, 5236—An address A P. C. A modern electric steel plant. K VON KESSELY, *Gussstahl-Ztg* 27, 101-2 (1930)—A description of the new elec. steel plant (wire, cable, etc.), at Campia Turzii, Roumania

CURTIS L. WILSON

Artificial atmospheres for electric furnaces and their application. A N OTIS *Iron Age* 126, 767-70, 844-8 (1930)—Results of Cu brazing of steel parts, of annealing sheet steel, strip and lamination punchings for elec. generators and transformers indicate that there is a wider field of application. A simple app has been developed for producing, from readily available materials, suitable gases rich in H at a fraction of the cost of H₂ produced by the usual methods. N may be used as a diluent of H. Where there is a demand for atomic H for welding, NH₃ may be dissociated by heat into a gas consisting of 75% H and 25% N. The production of *electrolene* is described. The H content prevents oxidation and the small CH₄ content prevents decarburization. The content of the gas is controlled by the operating temp of the dissociator. Several diagrams, tables and illustrations of different types of furnaces are included.

W H BOYNTON

Production of alloy steels in coreless induction furnaces. O DÖRRFENBERG AND N BROCKHO *Stahl u. Eisen* 50, 617-26 (1930)—After a review of the development of the high frequency induction furnace and a discussion of the elec. principles on which it is constructed, an account is given of the installation and performance of one of these furnaces in a German steelworks. The transformer efficiency is about 78-80%, and the current consumption for a 300-kg charge starting in the cold is 600 kw hr/ton to effect complete fusion and 730-790 kw hr/ton until casting is complete, the av. duration of melting is 90-100 min. Curves showing the elec. conditions during the melting of several charges are given, and the operation of the furnace is compared with that of a gas fired crucible furnace.

B C A

Tonnage melting by coreless induction. E F NORTHRUP *Iron Age* 127, 228-33 (1931)—The first of a series of articles on the theory, operation and uses of the high-frequency induction furnaces. For power requirements of 50 kw or less, the lig gap type of oscillator is best in present practice. For greater power, the motor or turbine driven generator giving frequency well within design and economic limits will best meet requirements. Mechanically considered, the steel housing is lined everywhere on its inner surface with high-cond. sheet Cu about 1/4 in. thick. This lining prevents eddy currents from being induced in the steel housing. The best results are obtained when using a "graded" coil without vacant spaces between turns.

G T, M

Valve-operated coreless induction furnace for high-temperature research. FRANK ADCOCK *Trans Faraday Soc* 26, 544-60 (1930)—A detailed description of the app. and circuit used is given, particular attention being paid to special features in design, safety appliances and the application of the furnace to metallurgical researches at high temps. Two air-cooled silica valves, each capable of dissipating 2.5-3.0 kw at the anode, are employed on the "push pull" or "back-to-back" principle. The valve filaments are rated at 40 amps and 17.2 v and are supplied with a c. by means of a transformer with a center tapped secondary. A combination of grid leak and condenser in conjunction with an accumulator bank maintaining a neg. potential of 36 v is used to apply the correct grid bias to the oscillating valves. Accurate regulation of the power supply is obtained by means of a condenser potential divider by means of which a definite fraction of the full oscillatory voltage in the furnace circuit is rectified and passed through a milliammeter used as an indicator. Frequencies of 50,000 and 1,000,000 cycles per sec. are used, special furnace construction (from the viewpoint of insulation between the turns of the inductor coil) being devised for the higher frequency. Owing to the high frequencies employed the furnace is capable of heating efficiently

charges consisting of small metallic fragments "Bridging" (i. e., the formation of a liquid pool by the lower portion of the charge while the upper portion remains as a relatively cool "bridge") of the charge on melting is not very pronounced at the 1 million cycles per sec. frequency. Detailed descriptions (with drawings and photographs) of the component parts of the furnace set-up are given. Observations on technique and applications of the furnace to the obtaining of high temp. thermal curves, and to magnetic and resistivity investigations are also included. EDWARD B. SANIGAR

Refining of ferrochrome in coreless induction furnaces. C. TAMA *Arch. Eisenhütte* 4, 55-61 (1930).—Production of Fe-Cr, decarbonization in coreless induction furnaces, suitable refractories for lining, and refining are considered. The systems Fe-Cr, Fe-Cr-C, and Cr-C are discussed in the light of the recent investigations. A charge of 300 kg. of Fe-Cr is usually melted in 70-100 min. in a 50-cycle furnace or in 60-80 min. in a 500-cycle furnace, and a charge of 30 kg. takes only 40 min. in a 10,000-cycle furnace. The power consumed varies from 750 to 1850 kw. according to the type of the furnace. G. T. MOROK

Chromium-plating of paper-mill rolls. R. E. CLEVELAND *Paper Mill* 54, No. 4, 14-5 (1931).—A brief discussion of the advantages of Cr plating paper mill rolls to increase their life by its hardness and resistance to corrosion. A. PAPINEAU COUTURE

Cathodic protection of metals in neutral solutions. U. R. EVANS, *Metals and Alloys* 2, 62-4 (1931).—A table gives values for i_c below which production fails for several steels and irons. A. J. MONACK

Buffer action in nickel-plating solutions. K. PITTSCHNER *Metal Ind. (N. Y.)* 29, 119-20 (1931).—The resistance to change in p_H exhibited by a soln. when it is subjected to gain or loss of acid or alkali varies greatly at different p_H values. Certain reagents, which are now common constituents in Ni plating solns., have the property of depressing the tendency toward large changes in p_H . The buffer effects of boric acid and of NH_4Cl on Ni content were studied. The extent of buffer action is evaluated by the Van Slyke method. It is the differential ratio of the increment in g equivs. of strong acid or base added per l of the soln. to the resultant increment or change of p_H . The buffer effect is highest at the upper end of the p_H range and least in the vicinity of p_H 4.0—only 1/20-1/100 of the value at p_H 6. In the usual plating range of p_H 6-5, the buffer effect is more noticeable when alkali is added, while in the range around p_H 2.5, the reverse is the case. The NH_4Cl content of Ni solns. is kept high enough so that the anode corrosion will keep the trend of p_H toward the upper limit of the range. Addns. of acid only are necessary. A brief discussion is included. Exptl. results are tabulated. W. H. BOYNTON

Studies in the electrodeposition of nickel. II. The effect of current density and temperature. J. B. O'SULLIVAN, *Trans. Faraday Soc.* 26, 533-9 (1930), cf. *C. A.* 24, 1801.—The influence of cathode $e. d.$ and temp. on the appearance, structure and deposition potential of Ni deposited from buffered $NiSO_4$ solns. having various p_H values has been studied. The solns. used and the prepn. of the deposits were as reported in Part I of this work (*C. A.* 24, 1801). Deposits were obtained at 17° and 35° at $e. d.$ s. of 2, 5, 10 and 15 ma./sq. cm. Deposition potentials were measured against a sat'd calomel half-cell, a capillary electrode tip in contact with the cathode surface being used. Tables showing the appearance of the deposits and tables of deposition potentials are given, as well as graphs of the deposition potentials against the logs of the $e. d.$ s. (giving straight lines). Tables of the deposition potentials at unit $e. d.$ (obtained by extrapolation of the above graphs) and of the slopes of the above graphs are also given. The varied surface structure of the Ni deposits is explained as due to variations in the throwing power (mainly dependent on the slope of the deposition potential- $e. d.$ curve) of the solns. The slope of the deposition potential log $e. d.$ curve and the deposition potential at unit $e. d.$ were found to vary from one soln. to another. From this it is deduced that the constants of the Freundlich adsorption isotherm are functions of the p_H of the soln. and of the nature of the buffering agent employed. III. Effect of small quantities of iron and aluminum. *Ibid.* 540-3.—In Part I of this work it was shown that the structure of a Ni deposit varies according to the p_H of the plating bath and the nature of the buffering agent employed, and it was suggested that this was due to the formation of some colloidal $Ni(OH)_2$ or basic salt in the cathode film and its co-pptn. with the metallic Ni. The investigation has been extended by examg. the effect of adding to the plating bath small quantities of inorg. salts which would be likely to yield colloidal hydroxides at a lower p_H than is required for $Ni(OH)_2$, i. e., at a p_H where $Ni(OH)_2$ would be exerting little influence. Fe (added as $FeSO_4$ or $FeCl_3$ soln.) and Al (as potash alum) were used for this purpose. Expts. were performed with solns. contg. 24 g. $NiSO_4 \cdot 7H_2O$ and 1.56 g. NaCl per 100 cc., with a Fe

Fe or Al to Ni ratio of 1/1000 or 1/200 at pH 's from 3.0 to 4.7, deposition being continued for 20 min at a c. d. of 10 ma./sq. cm. It was found that small amts of Fe salts had no appreciable effect on the electrodeposition of Ni, and it is suggested that this is due to the preferential deposition of Fe, whereby this metal is removed from the cathode film before it can form any colloidal compds. Al can accumulate in the cathode film until colloidal compds. are formed, and if this process proceeds far enough the structure of the deposit is altered, it becomes black or "burnt." The black deposits were found to contain appreciable amts of Al (0.0006 g. in 0.65 g. deposit) whereas none could be detected in the white deposits (which were somewhat finer-grained than those from solns. contg. no Al) so that it is suggested that the change is due to the co-pptn. of colloidal $Al(OH)_3$ with the Ni. A discussion is given of the possible ways in which colloidal hydroxide—either as a film at the cathode or, with an amphoteric colloid, as a film some little distance from the cathode—could cause finer grained deposits.

EDWARD B. SANICAR

Electrodeposition of zinc on aluminum from sulfate solutions. H. C. COCKS. *Trans. Faraday Soc.* 26, 517-26 (1930), cf. *C. A.* 24, 2058.—Measurements were made of the deposition potential of Zn from $N ZnSO_4$ solns. with first c. d. and secondly pH of the soln. as the only variable. Some static potentials were also detd. and observations made upon character of deposits, gas evolution and adhesion of bubbles. Expts. were done with buffered and with unbuffered solns. at c. d.'s of 4.5, 9, 18 and 26 amp./sq. ft. at a temp. of 20°. Com. Al sheet (contg. 0.14% Si and 0.58% Fe), prepd. by sandblasting at 5-10 lb. pressure, was used as cathode. Potentials were measured by the capillary tip method. Tables are given for both buffered and unbuffered solns. showing the effect on the deposition potential and on the deposits, of various addition agents and of pH . Graphs of the cathode potential against c. d. are also given. The potential measurements indicate that the addn. agents should increase the throwing power of an unbuffered $ZnSO_4$ soln. somewhat when its pH is 3-4. The order of merit of the addn. agents in giving smooth, fine-grained deposits was found to be gum arabic plus β -naphthol, gum arabic, β -naphthol, glucose, while a parallelism was found between the order of efficiency of the addn. agents in improving the deposits, increasing the change of deposition potential with c. d. and increasing the actual value of the polarization. The static potentials of Zn deposits and of sandblasted Zn sheet in both buffered and unbuffered solns. were found to vary with time and with pH . For a given soln. the deposition potential became more neg. with decrease of pH at a fixed c. d., and more neg. with an increase of c. d. at a fixed pH . These changes were barely detectable in the absence of addn. agents. Suggestions are given to explain these results. The deposits from the buffered and the unbuffered solns. were mostly dull white. The deposits improved with decrease of pH and increase of c. d. The addn. agents had marked effects on the deposits. Change of pH had a less marked effect on deposits from buffered than on those from unbuffered solns. Gas pits on the deposits usually increased with decrease of pH and increase of c. d. From agitated solns. there were generally no gas pits from solns. of pH 6 and only rarely from solns. of pH 5. A comparison of the pH values of the solns. employed obtained by the quinhydrone and by the colorimetric methods is given. Within the range of pH 3.0-6.0 the quinhydrone values were the lower (extreme deviations, 0.2 and 0.8 pH). A soln. contg. $N ZnSO_4$, 0.25 $N NaOAc$ and 1 g./l. of gum arabic, with or without 0.1 g./l. of β -naphthol, is suggested as likely to prove most suitable for Zn deposition on Al and its alloys.

E. B. S.

Electrolytic zinc. CHARLES W. CRYO. *Chem. Markets* 28, 253-6 (1931).—The advantages of the electrolytic over the smelter process are described.

E. H.

Adhesion of electroplated coatings. W. BLUM. *Metals and Alloys* 2, 57-9 (1931).—A review.

A. J. MONACK

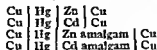
Electrodeposition of platinum, palladium and rhodium. W. KEITEL AND H. E. ZSCHNEGER. *Trans. Electrochem. Soc.* 59 (preprint) 4 pp. (1931).—The chloride plating bath for Pt metals is not satisfactory because of its short life. A new bath was accordingly developed in which the Pt is present as the diamminonitrite. Good, bright adherent deposits are obtained. The throwing power is very good. A similar bath for Pd was likewise developed and has given very satisfactory results.

C. G. F.

The limiting current density in the electrodeposition of noble metals. S. GLASTONK. *Trans. Electrochem. Soc.* 59 (preprint) 8 pp. (1931).—The literature of electrochemistry does not appear to contain a simple equation permitting the calcn. of the max. or limiting, c. d. at which a metal can be deposited with 100% current efficiency. Such an equation is now developed on the assumption that the max. rate of deposition is equal to the max. rate at which the ions contg. the metal to be deposited can diffuse up to the cathode. The values calcd. in this manner are shown to be in good agreement

with those observed, although discrepancies may arise when there is a possibility of H_2 evolution C. G. F.

Isothermal metallic cells. O. SCARPA. *Mem. accad. Italia, Classe sci. fis. mat. e nat. 1, Chim.* No. 5, 23 pages (1930).—This study deals with the degree of freedom of ions in solid and liquid metals. The motion of ions in metals, diffusion in solid metals, also the e. m. f. of contact of metals and thermodynamical calen. are considered in the first part of the paper. This cell may consist of the following:



With Zn amalgam a satn. point of 0.75×10^{-4} v. is reached at 19° with a concn. of 5.5 atomic per cent. of Zn G. T. MOTOR

Theory of electrolytic diaphragm cell. FRANCESCO GIORDANI. *Mem. accad. Italia, Classe sci. fis. mat. e nat. 1, Chim.* No. 6, 75 pages (1930).—A mathematical study of electrolysis based on the expts. carried out in a new type of electrolytic diaphragm cell for the production of caustic alkalis (developed by Giordani and Pomilio). This cell is set vertically with the graphite electrode in the center and the diaphragms on both sides of the electrode. The formula for instantaneous discharge in the electrolytic diaphragm cell can be written (A) $R = 1 - \{n/[1 + (x_1/x_2)]\}$ and the ratio x_1/x_2 can be expressed as a function of the concn. C_1 and C_2 of the chloride and the hydroxide dissociated at the cathode (B) $x_1/x_2 = Am(C_1/C_2)$ and (C) $C_1 = C_0 - BC_2$. A and B are functions of the initial concn. C_0 and temp. By substituting (B) and (C) in (A), there is obtained (D) $R = (1 + MC_2)/(1 + NC_2)$, where M and N are also functions of C_0 and temp. For higher values M is eliminated and (E) $R = 1/(1 + NC_2)$ results, the formula which was obtained by Guye G. T. MOTOR

The electrochemical oxidation of ketones. I. A preliminary paper. W. E. BRADY and CARL J. OPP. *Trans. Electrochem. Soc.* 59 (preprint) 8 pp. (1931).—Acetone was oxidized at a Pt wire gauze anode at low anodic e. ds. The course of the oxidation was followed by the collection and the analysis of consecutive portions of the gaseous oxidation products. The effect of variation of the e. d. on the compn. of the anode gas was investigated. The anode gas from the oxidation of acetone normally contained CO_2 , CO, O_2 , unsatd. hydrocarbons, C_2H_4 , and C_2H_2 . Analyses of the anode soln. were made. In every case, aldehydes, MeOH and pyruvic acid were absent. The anode soln. normally contained formic acid, acetic acid, mesityl oxide, unchanged acetone and a tar. At higher e. ds. phorone was identified. Efforts were made to show the disposition of all the acetone originally placed in the cell C. G. F.

The alleged electrochemical sulfonation of an aromatic hydrocarbon. FR. FICHTER, H. E. SUENDERHAUF and A. GOLDBACH. *Helv. Chim. Acta* 14, 249-53 (1931).—An emulsion of PhMe, EtOH and concd. H_2SO_4 was electrolyzed with an anode e. d. of 0.19 amp./cm². The alicidic portion of the product was converted into Ba salts. Ba-(OAc)₂ and (EtOSO₂)₂Ba $2H_2O$ (I) were the only products isolated. No CO_2 , $C_2H_5SO_3$, Ba

was found, contrary to reports in the literature (Puls, *Chem. Ztg.* 25, 263 (1901)). The electrolysis of a mixt. of EtOH and H_2SO_4 yielded the same products free from resins produced by the oxidation of the PhMe. On the other hand, a like operation on PhMe and H_2SO_4 yielded no acid. The electrochem. oxidation of acetophenone. H. E. SUENDERHAUF. *Ibid.* 14, 253.—The electrolysis of PhAc emulsified with H_2SO_4 led to the formation of small yields of maleic and fumaric acids. Some phenolic compds. were also present among the products I. M. LEVINE

Electrolytic reduction of 4-keto-3-phenyl-3,4-dihydroquinazoline. H. IROML. *Mem. Coll. Sci. Kyoto Imp. Univ.* 13A, 311-3 (1930).—Electrolytic reduction of 4-keto-3-phenyl-3,4-dihydroquinazoline with a lead cathode and aq. alc. Na_2CO_3 as the cathode soln. at 25° gives 4-hydroxy-3-phenyl-1,2,3,4-tetrahydroquinazoline, m. $170-1^\circ$, in 40% of the theoretical yield. Both the above quinazolines are reduced to 3-phenyl-1,2,3,4-tetrahydroquinazoline, m. $118-9^\circ$, when the cathode consists of Cu coated with Pt black and reduction is carried out at $50-60^\circ$. B. C. A.

Electrolytic Peltier heats and their measurement by isothermal, adiabatic, differential calorimetry. E. LANGE and J. MONHEIM. *Z. physik. Chem., Abt. A*, 150, 177-202 (1930).—The electrolytic Peltier heats at reversible electrodes are discussed in relation to other thermodynamical quantities characteristic of the electrolytic cell. Earlier measurements are reviewed, and an isothermal, adiabatic, differential calorimetric method is described and has been applied to the Hg | Hg⁺⁺ electrode, using a Hg elec-

trode in HgNO_3 soln and the normal HgCl electrode. The tabulated values show an increase in the pos. direction with increasing metal ion concn., in agreement with theoretical prediction. B. C. A.

Method of mapping equipotential lines and its application to electrical precipitator problems. A. W. SIMON AND L. C. KAOV. *Rev. Sci. Instruments* 1, 527-36 (1930). B. C. A.

Construction, operation and characteristics of photoelectric tubes. LEWIS R. KOLLER. *J. West Soc. Eng.* 36, 15-25 (1931).—A discussion of the characteristics of the different types of photoelec. tubes showing what kind of tubes should be collected with their operating limits for the different services. Metals of the alkali group are used in these tubes. Comps. of these metals change the sensitivities markedly. (One of the most useful characteristics of the photoelec. cells is the relation between current and illumination. This is strictly linear over a wide range. The photoelec. tube is sometimes called the "electric eye." Alone it is less sensitive than the human eye but with the aid of an amplifier it is 20 times as sensitive. Several types of amplifier are discussed. Where a great constancy in operation is required, vacuum tubes should be used. Outstanding applications of the tubes are television, talking pictures, smoke detectors, counters, bean sorters, light recorders, color analyzers, phys. photometers and innumerable control devices. W. H. BOYNTON.

Hydrogen-cooling for turbine-generators. M. D. ROSS. *Elec. Eng.* 50, 211-4 (1931).—Actual operating experiences with a 7500-kv amp unit are outlined. Five advantages of H_2 as a cooling medium for rotating elec. machinery are: (1) Windage losses are reduced to about 18% of their value in air. (2) For a given amt. of active material the rating of a generator with H_2 cooling is 25% greater. (3) Corona has little effect on insulation in H_2 atm., increasing the life of the insulation. (4) Fires in the generator are impossible, on account of absence of O_2 . (5) Smaller gas coolers and less cooling water are required than with air cooling. W. H. BOYNTON.

The relative merits of gas, oil and electricity for industrial purposes (HOPKINSON) 13. Special refractories for electric furnace linings (KUKLA) 19. Some new facts of a chemical nature in the field of corrosion investigations and in the protection of metallurgical products against corrosion (COURNOT) 9. Corrosion of steel water pipes by stray electric currents (ROTHIE) 9. Electrochemical oxidation of paraffin (ATANASIU) 22. Au and Ag solutions (Brit. pat. 338,383) 18. Heat treatment of loaded electrical conductors (Brit. pat. 338,160) 9. Rubber battery boxes (Brit. pat. 338,114) 30.

BILLITER, JEAN. *Électrometallurgie des solutions aqueuses*. Translated from 2nd German ed. by J. Salaure and S. Salaure. Paris: Dunod, 324 pp. 184.

Dry batteries. SOC. ANON. LE CARBONE. Ger. 514,966, Feb. 16, 1924. The Zn and C electrodes are embedded in an electrolyte mixed with peccunizable colloid and powd. C. The example mentions a mixt. of NH_4Cl , arrowroot starch and powd. wood charcoal.

Dry cell electric battery. H. A. BUMKE-GES. Brit. 337,869, Aug. 27, 1929. Mech. features.

Storage battery. SOC. ANON. DES ACCUMULATEURS MONOPLAQUE. Brit. 337,375, Feb. 5, 1929. Between each pair of neg. plates there are placed two pos. plates each having half the thickness of the neg. plates, the pos. plates being joined together and enclosing electrolyte which tends to remain at max. concn. The electrolyte passes through the pos. plates under the influence of hydrostatic and electro-osmotic pressures. Alternatively, a single hollow pos. plate contg. an inert packing may be used.

Galvanic batteries. KARL KUPKA. Ger. 514,967, May 23, 1928. Details of arrangement of the electrodes in multicell batteries are given.

Selenium cells. FIRMA CARL ZEISS. Ger. 514,911, Jan. 8, 1929. Se cells contg. two Pt electrodes mounted on glass or quartz and a channel contg. Se are described.

Selenium cell of the condenser type. "SELENOPHON" LICHT UND TONBILDESSEL-SCHAF. M. B. H. Ger. 514,972, Feb. 4, 1930. Details of arrangements are given.

Light-sensitive selenium cells. TELEFUNKEN-GES. FÜR DRAHTLOSE TELEGRAPHIE. Brit. 337,691, Feb. 13, 1929. A cell formed on polished glass comprises electrodes formed by graphite deposited from a colloidal suspension on portions previously etched the suspending medium being driven off by heating and the interstices coated with Se or Se and Te together.

Electrodeposition of metals on rotating cathodes. S. O. COWPER-COLES. Brit.

338,173, May 16, 1929 In the deposition of metals such as Cu, Zn, Ni or Ag, a slowly rotating cathode is used which is supported on rollers above an arc-shaped anode and driven by rollers. The anodes may either be formed of the metal to be deposited or of inert material such as Pb, C or ferro-Si and the electrolyte may contain the compd. of the metal to be deposited either in soln. or in suspension. Deposition may be effected directly from ore. App. is described.

Forming acoustic diaphragms of thin metal such as nickel by electrodeposition. ALBERTIS HEWITT (to Victor Talking Machine Co.) U S 1,793,483, Feb. 24. Various details of app. and procedure are described.

Nickel-coated articles. FRANK V. KNAUSS (to Madsenell Corp.) U S 1,793,036, Feb. 24. A Ni surfaced article such as a spring steel bar is treated as anode in a H_2SO_4 bath and a fresh Ni coating is then electrodeposited on the treated Ni surface. A firmly adherent coating is formed which takes a good polish.

Electroplating ferrous metal articles with zinc. SIDNEY H. DAVIS, CARL O. ANDERSON, RUDOLPH J. STENGL, WM. N. SMITH and HERBERT R. HANLEY (to Century Zinc Co.) U S 1,795,081, March 3. Articles are first provided with a thin coating of Zn by electrodeposition from a ZnSO_4 soln. of suitable acidity with a current of suitable d., the articles are then transferred to a second bath of ZnSO_4 soln. of relatively higher acidity and the plating is continued with a current of higher d., and the c. d. is finally lowered and the plating further continued to produce a coating of the desired thickness.

Electroplating zinc on iron or steel. SIDNEY H. DAVIS, CARL O. ANDERSON, WM. N. SMITH and HERBERT R. HANLEY (to Century Zinc Co.) U S 1,795,079, March 3. A thin initial coating is formed on an article by current from an insol. anode in a bath comprising a ZnSO_4 soln. of relatively low acidity by using a d. c. of relatively high c. d., and further plating is effected with a current of lower d. in a bath of higher acidity with further use of an insol. anode. U S 1,795,080 relates to plating with Zn after preliminary deposit of a thin coating of spongy Pb from an all. Pb-bearing soln. by current of low d., from an insol. anode and brushing to remove some of the Pb and cause penetration of the ferrous metal by the remaining Pb.

Electroplating apparatus suitable for continuous operation. CHRISTIAN H. JORDING (to John W. Brown Mfg. Co.) U S 1,793,551, Feb. 24. Structural features.

Anode container for electroplating apparatus. THOMAS G. MELISH U S 1,792,998, Feb. 17. Structural features.

Apparatus for electroplating successive portions of a large area such as the interior of oil stills, etc. KEVIN W. SCHWARTZ (to United Chromium, Inc.) U S 1,794,497, March 3. Various structural details are described including means for moving the anode to different positions for use.

Apparatus for electroplating articles such as in the manufacture of "copper-clad" roofing. CHARLES E. YATES (to Anaconda Sales Co.) U S 1,794,748, March 3. Structural features.

Plating interior surfaces of metallic vessels. EDWARD O. DUNKLEY (to Standard Oil Co. of Calif.) U S 1,793,069, Feb. 17. Various details of app. and procedure are described for operations such as plating the interior of large metal vessels with Cr.

Continuous process for chromium-plating metal wires or strips. BYRON V. McBRIDE (to Westinghouse Elec. & Mfg. Co.) U S 1,794,973, March 3. The wire or strip is drawn downward through a Cr-plating soln. in such a manner that all portions are substantially equidistant from the anode and exposed directly to the voltaic action of the bath, and then drawn upward through the soln. in such a manner that all portions are again substantially equidistant from the anode and exposed directly to the voltaic action of the bath. App. is described.

Selectively chromium-plating portions of surfaces such as those of reflectors. VICTOR L. SODERBERG U S 1,794,929, March 3. Portions of the surface which are to remain unplated are covered with nitrocellulose lacquer preliminarily to the plating operation.

Electrolytic cell. I. G. FARBENIAND A. G. (Georg Pfeleiderer, inventor) Ger. 514,741, Nov. 2, 1926. Addn. to 471,925 (C. A. 23, 2376). In a bipolar cell, especially for the decompn. of water (as described in 471,925), a portion of the electrolyte is passed continuously through an outside circulatory system to regulate the temp. and concn. The app. is described.

Electrolytic cell. SOCIÉTÉ D'ÉTUDES POUR LA FABRICATION ET L'EMPLOI DES ENGRAIS CHIMIQUES. Ger. 514,501, Oct. 25, 1928. The cell has a middle partition and porous containers between the partition and the electrodes, the containers being raised from the bottom of the cell to permit free flow of the electrolyte.

Construction for maintaining a constant electrolyte level in electrolytic cells. CARL ROTH (to I G Farbenind A-G) U S 1,793,130, Feb 17.

Electrolytic production of compounds containing active oxygen. JOHANNES FCKELL (to I G Farbenind A-G) U S 1,792,633, Feb 17. See Brit 313,124 (C A 24, 792)

Metal hydroxides by electrolysis. R S CARREAS Brit 338 178, July 9, 1929 Hydroxides of Zn, Cu, Sb and other metals (other than Bi hydroxide and white lead) are obtained by use of an anode of the metal the hydroxide of which is to be formed, in a dil soln of an electrolyte such as NaClO_4 or KClO_4 satd with CO_2 (the electrolyte being continuously withdrawn from the cell without removal of the pptd hydroxide and returned after resatn with CO_2) Various details of app and procedure are described

Apparatus for electrolyzing fused alkali metal halides. DEUTSCHE GOLD- UND SILBER SCHEIDANSTALT VOM ROSSLER Ger 517,256, May 15, 1924

Apparatus for preparation of light metals by electrolysis of their fused halogen salts. SOC ANON POUR L'IND CHIM A BALE SWISS 142 518, Aug 23, 1929

Formation of gang by the decomposition of electrolytic solutions within the rock layers. SIEMENS-BALUNION G M B H, KOWM-GE8 (Michael Muller, inventor). Ger 514 818, Jan 17, 1929 Rocks not contg a natural electrolyte are impregnated with Na_2SiO_3 so that SiO_2 is freed on electrolysis The process can be used for petrification.

Aluminum production. COMPAGNIE DE PRODUITS CHIMIQUES ET ELECTROMETALLURGIQUES ALAIS, FROGES ET CAMARGUE Brit 338,044, March 15, 1929 A mixt of AlF_3 and Na or K carbonate or hydroxide is used instead of cryolite as a flux in the electrolytic reduction of alumina Various details, proportions, etc., are given

Protective layers of lead peroxide. SIEMENS & HALSKE A-G (Walter Birett and Johannes Fischer, inventors) Ger 514,621, Oct. 23, 1927 Anode conductors are given a protective coating of PbO_2 by electrolysis in an alk soln of a Pb compd contg also such O-contg org compds as aldehydes, ketones, alcs or their derivs Examples mention baths contg NaOH , hydroquinone and PbO , or NaOH , acetone, AcCl and PbO , or NaOH , fruit sugar and PbO

Bleaching shellac. OLGA MYLO NEE ROSENHAGEN Ger 517,000, May 12, 1929 A soln of shellac is treated with a soln of a suitable salt, e g, NaCl , and the mixt. is electrolyzed with stirring and cooling Preferably, the electrolysis is begun with a weak current, which is afterward strengthened An automatically operating app is described

Treating hydrocarbons in an electric arc. I G FARBENIND A-G (Otto Eisenhut, inventor) Ger 514,592, Dec. 15, 1925 App for subjecting satd hydrocarbons, e g, CH_4 , to the elec. arc to produce unsatd hydrocarbons, e g, C_6H_6 , is described Cf C A 24, 1586

Acetylene and hydrogen production in the electric arc. OTTO EISENHUT (to I G Farbenind A-G). U S 1,794,004, Feb 24 See Fr 674,459 (C A 24, 2385).

Electric furnace. DET NORSKE AKTIESELSKAB FOR ELEKTROKEMISK INDUSTRIE Ger 514, 675, May 7, 1926 Details of the electrodes are given

Electric induction furnace. EDWIN F. NORTHRUP (to Ajax Electrothermic Corp) U. S 1,795,136, March 3 Elec features

Electric induction furnaces. SIEMENS-SCHUCKERTWERKE A-G (Rudolf G Bertbold, inventor) Ger 517,348, July 17, 1926 Details of the current feed are described

Electric resistance furnace. HERBERT J MCCAULEY. U S 1,794,310, Feb 24 Structural features

Electric arc furnace for alloying or refining metals EMILIE BORNAND and HANS A SCHLAEFFER Ger 517,285, Mar 1, 1923

Electric annealing furnace. AKT GES BROWN, BOVERI & CIE Swiss 142,004, Sept. 17, 1929

Electric furnace for smelting metals. EMIL F RUSS U S 1,793,137, Feb 17. Structural features

Electric-furnace reduction of alumina. VEREINIGTE ALUMINIUMWERKE A-G Brit 337,995, Nov 27, 1928 A furnace is continuously charged automatically in accord with the quantity of alumina which is reduced to metal in the bath, the charging hopper being mounted on a hollow anode and discharge from it being controlled by a current indicating device.

Closed tiltable electric furnace suitable for iron and steel production EMIL G T. GUSTAFSSON and BENGT I NOREN (to Hampus G E Cornelius) U S 1,794,455, March 3 Structural features

Melting finely divided dry materials such as iron oxide in an electric induction furnace. LOWIN F. NORTHUP (to Ajax Electrothermic Corp.) U. S. 1,794,863, March 3. Lumps of cond. material such as pure iron which are large enough to be separately heated by induction in passing an a. c. about the charge are mixed with the charge.

Producing carbide-forming metals such as titanium, vanadium, chromium, iron or manganese. TURE R. HAGLUND U. S. 1,794,401, March 3. Ores such as oxides of the metals named are fused in an elec. furnace together with added reducing agent including lumps of material contg. bauxite, a basic refractory oxide of lower m. p. than the ore, such as CaO, and carbonaceous reducing material. A cement may be produced as a by-product.

Electric furnace with carbide resistances. R. C. BENNER and G. J. EASTER (to Carborundum Co. and C. I. Hawke) Brit. 338,131, Dec. 7, 1928. Various structural details are described of furnaces which may have resistances of Si carbide or B carbide surrounded by a body of granular carbon such as crushed petroleum coke.

Electromagnetic stirring device for furnaces for the electrolysis of fused aluminum salts. HANS THOMA. Ger. 514,919, Dec. 4, 1928.

Device for electrically heating and for spraying molten metals for coating surfaces. WM. S. RICH and CHARLES M. SAEGGER, JR. U. S. 1,792,551, Feb. 17. Structural features.

Electrode mounting for electric furnaces or electrolytic cells. A. FRANCHINI (to P. Girod) Brit. 338,371, Aug. 27, 1929. Structural features.

Lattice cathodes. RAGNAR ANHALTER METALLSCHERREI N. D. H. Ger. 514,716, June 21, 1929. Lattice cathodes for the electrolysis of alkali chlorides are described.

Coating cathodes with alkaline earth metal oxides. FRITZ TÖDT. Ger. 514,600, Apr. 8, 1928. The glowing cathode is coated with alk. earth metal oxides, especially NaO, by electrolysis in an aq. bath of alk. earth metal salt contg. CO₂ ions and heating the carbonate deposit so formed to glowing.

Cathode for rotary flame arcs. GEORGES GAUTIER, LEON LE BOZEC and RAYMOND DUBOIS. U. S. 1,793,605, Feb. 21. A cathode comprises a hollow annular ring in which a cooling fluid flows continuously and a winding immersed in the fluid which is supplied with elec. current and has its axis the same as that of the ring.

Glowing cathodes. SIGMUND A. G. Swiss 142,281, Aug. 6, 1929. Glowing cathodes are made with a core of non readily fusible metals. This is heated with one or more alkali metals in the absence of O to form a surface alloy, which is then oxidized to complete the cathode.

Incandescent cathode for electric discharge tubes. ANTON E. VAN ARKEL and JOHANNES BRUIJNES (to N.-V. Philips' Gloeilampenfabrieken) U. S. 1,794,810, March 3. Cathodes are formed with a metallic core such as Ni, W, Mo or Pt and a surface consisting of a nitride of at least one of the elements Ti, Zr and Hf.

Luminous electrical discharge tube and associated cooling device. JACQUES RISLER (to Risler Corp. of America) U. S. 1,793,720, Feb. 24. Structural features.

Electric gas-cleaners. METALLGES. A. G. (Heinrich B. Röder, inventor). Ger. 517,261, July 18, 1929. Means is described for keeping the high tension insulators dry by supplying radiant heat.

Electric gas-cleaners. METALLGES. A. G. (Kurt von Lberhard, inventor). Ger. 517,262, Oct. 5, 1929. Details of the discharge electrode and its mountings are described.

Electric gas-cleaning apparatus with parallel-plate precipitating electrodes. METALLGES. A. G. (Jakobus R. Gies, inventor). Ger. 517,067, June 8, 1928.

Hammer device for gas-cleaning electrode. SIEMENS SCHUCKERTWERKE A. G. (Alexander Kaufmann, inventor). Ger. 517,263, June 26, 1928.

Electric gas-purification plant. METALLGESSELLSCHAFT A. G. Ger. 514,587, Dec. 9, 1928. Addn. to 511,643 (C. A. 25, 1169). Details of cleaning the electrodes by shaking.

Electric gas-purification plant. SIEMENS SCHUCKERTWERKE A. G. (Richard Heinrich, inventor). Ger. 514,650, July 31, 1927. Details of the pptg. electrodes are given.

Chamber for the electric purification of gas. SIEMENS-SCHUCKERTWERKE A. G. (Richard Heinrich, inventor). Ger. 514,820, Mar. 10, 1926. The pptg. surfaces are formed of tin plates.

Apparatus for electrical precipitation. CARL W. J. HEDBERG and HARRY A. WINTERMUTH (to Research Corp.). U. S. 1,794,074, Feb. 24. Structural features.

Apparatus for electrical precipitation of suspended particles from gases. LODGE-COTTRELL, LTD., and L. LODGE. Brit. 337,642, Nov. 29, 1929. Structural features.

Apparatus for electrical precipitation of suspended particles from gases. Me-

positive induced dichroism is obtained on development. The polarized initial exposure with the CrO_3 treatment had changed the sensitive film to a photochem polarizer. The effect is a support for the micelle theory of latent image and was predictable from it. It is expected that the same Albert reversal could also be obtained with highly sensitive, coarse grained emulsions.

B. J. C. VAN DER HOEVEN

Determination of the general and color sensitivity of negative materials. C. EMMERMANN. *Schweiz. Photo Ztg.* 32, 220-2, 224-9(1930).—The principles and the application of the methods of Scheiner, of Fiedler-Recht and of Hurter and Driffield are discussed for the determination of the general and color sensitivity of neg. materials. H. D. RUSSELL

Optical sensitizing of silver halide by colloidal silver. S. B. SHEPPARD. *J. Franklin Inst.* 210, 587-607(1930).—Various theories on the optical sensitizing of Ag halides are discussed. It is suggested that optical sensitizing by colloidal Ag (Becquerel effect) is primarily due to sensitized photoelec. emission by Ag amicros and ultramicros. Sensitized photoelec. emission of composite metal films is considered as due to minute Volta condenser fields. Self fields as origins of sensitized photoelec. effects are considered in relation to the Becquerel and Weigert effects with Ag halides. E. P. WIGHTMAN

Alteration of orthochromatism by pinakryptol. K. HAIDRICH. *Photofreund* 11, 50-2(1931).—Sensitometric tests were made on a series of plates. The same exposure was given each plate. One plate was desensitized with a normal pinakryptol bath before development. A normal developer (1/20) was used. The normal plate was developed for 5 min. and the desensitized plate for 6 min. in total darkness. With the Gevaert SSS plate a loss of 5° Scheiner resulted, with Gevaert Super-Press 8° Scheiner; with Agfa Isochrome 9° Scheiner and with Agfa Alkocyanine plates 15° Scheiner. It was also found that desensitizing affected the color sensitiveness of an emulsion as follows:

Agfa Isochrome plates		
Blue	green	yellow
3° S	3° S	3° S loss in speed
Gevaert Super-Press plate		
Blue	green	yellow
6° S	8° S	8° S loss in speed
Gevaert SSS plate		
Blue	green	yellow
1° S.	4° S	3° S loss in speed
Alkocyanine plate		
Blue	green	yellow
14° S	11° S.	16° S.
		red
		16° S

loss in speed

increase in speed

A. K. WITTMER

Desensitizing. NEUGEBAUER. *Photofreund* 10, 300-3(1930).—Sensitometric tests were made to det. loss of speed by the use of a desensitizer (pinakryptol green as a preliminary bath). A desensitized plate compared with untreated plate showed a loss of speed when both were given equal development. If the developing time is doubled, an increase of speed is obtained which approaches but does not usually equal that of the untreated. If a desensitizer is used two to three times the normal exposure should be given.

A. K. WITTMER

Desensitizing. H. DÖRR. *Photofreund* 10, 441-3(1930); cf. Neugebauer (preceding abstract).—D. directs attention to several publications by O. Mente, A. Hubl, K. Jacobsohn, M. L. Dundon, J. I. Crabtree and B. H. Carroll on desensitizing. All authors claim no loss of speed if a plate or film is developed a little longer than an ordinary plate which is not given a desensitizing treatment.

A. K. WITTMER

Desensitizing. LÖFFO CRAMER. *Photofreund* 11, 7-9(1931).—Exception is taken to the view of Neugebauer (2nd preceding abstract) that the use of desensitizers results in a loss of speed. L. C. cites some of his own papers in support of his statements.

A. K. WITTMER

Desensitizing. NEUGEBAUER. *Photofreund* 11, 27(1931).—In reply to Löffo-Cramer (preceding abstract) N. states that the newer emulsions may act differently with a desensitizer from emulsions made in 1920-1925.

A. K. WITTMER

Chromium and sulfur intensification of negatives and prints. H. D. POWER. *Camera Craft* 38, 73(1931).—The method consists in bleaching the negatives or prints in a soln. of 2% $\text{K}_2\text{Cr}_2\text{O}_7$ plus 1.5% strong HCl and substituting a freshly prepd. 2% soln. of Na_2S for the developer. The use of a bisulfite soln. is mentioned to discharge the color of the dichromate to give satisfactory results.

G. W. WILHELM

Some properties of fine-grain developers for motion picture films. H. C. CARTON AND J. I. CRANTON. *Trans Soc Motion Picture Eng* 13, 408-444 (1929). Graininess can be lessened by (1) increasing the sulfite, (2) decreasing the rate of development and (3) increasing the degree of development. Borax developers do not give aerial fog, but also do not produce images with less graininess. Tables showing the effect of compo, agitation, aging and exhaustion and time fog curves are given. A. N. H.

Borax developer characteristics. H. Meyer AND D. R. WINTER. *Trans Soc Motion Picture Eng* 13, 445-52 (1929). A sulfite content of 2% per l. gives grain free images with high speed. The effect of borax is due to its alkali. Metol rather than hydroquinone does the developing. Hydroquinone tends to increase fog. Renewing a borax developing bath brings back the development rate but not the detail giving power which has been lost through bromide accumulation. ANN. NICHOLSON 1180.

Neutral and acid amidol developer. C. FIMMERMANN. *Phot. Chronik* 37, 134-40 (1930). Acid amidol is a somewhat slower developer than neutral amidol. Emulsion speed is not affected. Acid amidol also produces more fog. Halation is more noticeable with the acid amidol. A. M. ROBINSON.

Warm tone developing, selenium and sulfur toning. C. FIMMERMANN. *Phot. Chronik* 38, 10, 15-27 (1931). The use of pyrocatechol, glycine, hydroquinone and kigonogen without sulfite as warm tone developers is lately discussed. 'a' and 'B' toning are also briefly described. A. M. ROBINSON.

"Hyalotype" and its paper prints in natural colors. E. WINTER. *Cameras (Luzern)* 9, 185-190 (1911). Three color sepia negatives are made in a repeating back camera. A print is made from each of these on dichromated gelatin film by exposure through the support. A wash off relief is then produced on each film. The reliefs are dyed in the three complementary colors and a three color multiplication print is made upon gelatin coated paper. This multiplication process remedies the poor shadow quality and variation in successive prints characteristic of other multiplication processes. This end is achieved by rinsing the dyed reliefs in dist. water before making the prints and by employing dye solutions of the correct degree of dilution. Diffusion of the dye in the printing paper is prevented by incorporating casein into the gelatin. M. W. B.

Positive material for motion pictures in color. R. LANDAU. *Kino-technik* 13, 12- (1931). A hypothetical color printing process is described in which 3 fine grained photographic emulsions are to be made, each of a million grain size differing from the other two. When developed in the same developer, each emulsion is to give a colloidal Ag deposit of one of the 3 complementary printing colors. The emulsions are to be coated in superposition upon a common film support. The top and bottom emulsions are not to be color sensitized, while the middle one is to be orthochromatic or panchromatic. All are to be impregnated with a yellow screening dye. The lowest is to give a red Ag deposit, the middle one a yellow deposit and the top one a blue deposit. The film is exposed 3 times, each time through one of the three color sepia negatives. In the green filter neg., through the support, with blue light, in the red filter neg., from the emulsion side, with blue light, in the blue filter neg., with yellow light. The film is then to be developed to obtain a pos. in 3 colors. Instead of making the emulsions of 3 grain sizes, they may be of one grain size, and the developer, preferably containing a Ag halide solvent, may be altered in compo. as it penetrates the successive emulsions. M. W. BAYMOR.

Photographic detection of the asymmetrical angular distribution of twice reflected electrons (RUFF) 3. Direct photography of ionization in ionizing substances (GRIMANT) 3. Method for preparing photographs of Petri dish cultures by direct contact printing on photographic paper (HUCHMAN, L. W.) 11C.

NEWENS, FRANK R. *The Technique of Colour Photography*. London. Blackie and Son, Ltd. 101 pp., 4s. 6d., net.

Photographic emulsions. H. H. SHERRARD AND P. P. WIGHTMAN (to Kodak, Ltd.). Brit. 317,121, Feb. 25, 1929. Stability of emulsions is improved and any tendency to fog is corrected by adding small proportions of compounds such as saccharin, diacetamide, phthalimide, pyridine derivatives such as acetylaminide and ring compounds such as those containing pyrazoline, pyrazide, pyrazolone or s-triazole ring. (I. C. A., 24, 3720).

Photosensitive material. H. D. MURRAY, D. A. SPENCER AND COLUNA PHOTOGRAPHIC (HARRIS & PIERSON), LTD. Brit. 317,218, Aug. 23, 1929. A water impermeable cellulose ester such as a nitrocellulose or cellulose acetate compo., or bitumens

material such as paper or fabric coated or impregnated with such an ester, has its surface rendered water permeable by superficial de-esterification and is then impregnated with a light-sensitive salt such as a compd. of Ag, Cr or Fe or a diazo compd. Various details of treatment are described.

Apparatus for developing photographic materials such as sensitive paper by a developing gas such as ammonia. ADOLPH LANGSTRA (to Eugene Dietzgen Co.). U. S. 1,794,603, March 3. Structural features.

Bromide printing. OTTO TARNICHEL. Ger. 493,920, Sept. 26, 1928. Bromide prints without spots or wrinkles are produced by having at most 1.8% alkali and at least 1.8% Na_2SO_4 in the KHSO_4 and citric acid developer. A suitable developer contains metol, hydroquinone, Na_2SO_4 , Na_2CO_3 , KHSO_4 , citric acid and KBr .

Blue prints. KALLE & Co. A-G (Maximilian P. Schmidt and Rudolf Zahn, inventors). Ger. 517,159, July 24, 1926. A compd. (other than a dye) showing strong absorption in the ultra violet, e. g., Na 6-phenyl 2-aminopseudoaminobenzene-3,4'-disulfonate, is added to inks or crayons for making drawings from which blue prints are to be prep'd. Clearer reproduction results.

Screen for use in photoengraving. JOSEF W. LIPPERS (one half to Joseph Deutsch). U. S. 1,794,603, March 3. In forming a patterned, irregularly reticulated, hardened collod screen for photoengraving, a dichromated collod layer is exposed to actinic rays at predet. areas and again exposed to light in its entirety, and swelling is then effected by development with suitable collod-swelling liquids, followed by washing and drying.

Photomechanical printing surfaces. EUSTACE B. ELDREDGE (one half to John A. Haeseler). U. S. 1,793,070, Feb. 17. For producing a substantially grainless printing surface from a latent Ag image in a colloidal material, capable of being printed from as in collotype printing, the exposed plate or film carrying the latent image is developed with a developer such as a dil. amidol soln. which reacts upon all the light affected Ag salt in the collod proportionately to the amount of light to which the Ag salt has been exposed, but the action of which is so feeble as to produce an extremely faint and tenuous image lacking d. and depth which will be composed of very minute quantities of reduced Ag held in suspension in the collod, and the collod is chem. hardened in the presence of the Ag (suitably by the action of an aq. soln. formed from CuSO_4 , KBr , $\text{K}_2\text{Cr}_2\text{O}_7$ and chromic acid).

6-INORGANIC CHEMISTRY

A. E. MIDDLETON

The atomic weight of iodine. The analysis of iodine pentoxide. GREGORY P. BAXTER AND ALBERT Q. BUTLER. *J. Am. Chem. Soc.* 53, 169-77 (1931).— I_2O_5 produced by dehydration of HIO_3 contains less I (and more O) than corresponds to the theoretical formula. I_2O_5 of normal compn. is difficult, if not impossible, to prep. Conclusions based upon its actual compn. are not sufficiently reliable for work of high precision, as attested by the at. wt. obtained, 126.905. B. A. SOTTE.

Principles of the structure of boron-hydrogen compounds. Addendum. E. WIBERG. *Z. anorg. allgem. Chem.* 195, 258 (1931).—Because of an abbreviation of ms. (cf. C. A. 25, 1751) omission was made from the table of properties of B_2H_6 and C_2H_2 , that the statement regarding addn. of Na to the ethylenic bond relates to derivs. of ethylene and that the assumption of addn. of hydrogen halides, halogens and compds. of the type HOX to B_2H_6 relates to a previous hypothesis on boron hydride substitutions elaborated by W. A. P. SACHS.

The chemistry and metallurgy of beryllium. G. MALCOLM DYSON. *Chem. Age (London)* 24, 228-30 (1931).—A review. E. H.

The separation of beryllium and aluminum oxides. A. TRAVERS AND SCHNOUTKA. *Compt. rend.* 192, 295-7 (1931).—Methods available are based on these facts: (1) BeO dissolves in a soln. of NH_4HCO_3 while Al_2O_3 does not. (2) Solubilities in alk. solns. are different. (3) Solubilities in a boiling soln. of NaHCO_3 are different. (4) A basic acetate of Be is sol. in CHCl_3 . (5) Al is pptd. by 8-hydroxyquinoline. T. and S. offer the following methods: (1) The bisulfates are evapd. to dryness on a steam bath and exhd. Al forms an invol. oxide but Be does not. (2) $\text{Be}(\text{HSO}_4)_2$ ppts. in boiling soln. in the presence of traces of H_2SO_4 but redissolves on cooling, while $\text{Al}(\text{HSO}_4)_3$ does not. (3) Be forms a complex with NaHSO_4 which is completely decomposed by boiling and Be does not form compds. with Al at pH 4. An extensive bibliography is given. A. L. T.

Polytherms of ternary systems which in addition to water contain an alkaline sul-

fate and manganous sulfate. III. A. BENRATH. *Z anorg allgem Chem.* 195, 247-54 (1931), cf. *C. A.* 24, 3159.—The polytherms are detd for the systems $\text{MnSO}_4\text{-Ti}_2\text{SO}_4\text{-H}_2\text{O}$, $\text{MnSO}_4\text{-(NH}_4)_2\text{SO}_4\text{-H}_2\text{O}$, $\text{MnSO}_4\text{-Rb}_2\text{SO}_4\text{-H}_2\text{O}$. The solid phases present are $\text{MnSO}_4\cdot\text{Ti}_2\text{SO}_4\cdot 6\text{ aq}$, $2\text{MnSO}_4\cdot\text{Ti}_2\text{SO}_4\cdot\text{Ti}_2\text{SO}_4$ and corresponding salts of NH_4 and Rb , in addn to the mono-, penta- and heptahydrate of MnSO_4 in each system. The temp range was 0-100°. K does not form the double salt $\text{MnSO}_4\cdot\text{K}_2\text{SO}_4\cdot 6\text{ aq}$, while with Cs the salt $2\text{MnSO}_4\cdot\text{Cs}_2\text{SO}_4$ is lacking.

II. STOERTZ

Ternary systems. IX. Sodium iodate, sodium nitrate and water. ARTHUR E. HILL AND JOHN E. DONOVAN. *J Am Chem Soc* 53, 934-41 (1931), cf. *C. A.* 23, 4616.—The soly of NaIO_3 in water is given for temps from 5° to 50°. Isotherms in the ternary system $\text{NaIO}_3\text{-NaNO}_3\text{-H}_2\text{O}$ are given at 5°, 25° and 50°. The stable hydrates are $\text{NaIO}_3\cdot 5\text{H}_2\text{O}$ below 20° and $\text{NaIO}_3\cdot\text{H}_2\text{O}$ above this temp. In the ternary system, the double salt, $2\text{NaIO}_3\cdot 3\text{NaNO}_3\cdot 15\text{H}_2\text{O}$, appears below 9.8°. X. Magnesium iodate, magnesium nitrate and water. ARTHUR E. HILL AND SAMUEL MOSKOWITZ. *Ibid* 941-6.—The soly of MgIO_3 in water is given from the eutectic (-0.36°) to +90°. In addn to the deca and tetrahydrates, previously known, the anhydrous form is shown to exist in contact with water above 57.5°. Isotherms in the ternary system $\text{Mg(IO}_3)_2\text{-Mg(NO}_3)_2\text{-H}_2\text{O}$ are given at 5°, 25° and 50°. No double salts are found in this temp range.

B. E. TIFFANI

Behavior of sodium sulfate crystals toward cold ethyl alcohol and methanol. M. A. RAEVSKY. *Z Krist* 73, 115-6 (1930)—Treatment with 95% EtOH affords a powder consisting of thenardite, with MeOH dehydration, but not powder formation, occurs.

B. C. A.

The chlorides of sulfur, T. M. LOWRY. *Z Electrochem* 36, 733-4 (1930), cf. *C. A.* 24, 441.—Dielec consts for mixts of S_2Cl_2 and Cl_2 in the solid state show the presence of S_2Cl_4 . $F\text{-p}$ curves indicate the sepn of solid S_2Cl_4 , SCl_2 and SCl_4 . Surface tension, d , light absorption and parachors show the liquid mixts behave as a ternary system $\text{S}_2\text{Cl}_2\text{-SCl}_2\text{-Cl}_2$.

O. M. SMITH

Sulfur fluoride, KÄRLLS STRENGS. *Acta Univ Lohrens Ksm Fakultat Seriya f*, 233-60 (in German 261-2) (1930)—The interaction of AgF and S yields Ag_2S and S_2F_4 , a colorless gas with an odor similar to that of SO_2 or S_2Cl_2 , it fumes in air and is very toxic to the organism, it is decomposed by water, yielding S , H_2SO_4 and HF , it attacks glass, even at low temps, and yields S , SiF_4 , SOF_2 and SO_2 , its formula has been proved by analysis and density detn. It attacks quartz, and consequently the physical constants reported by Fischer and Jaenckner (*C. A.* 23, 5014) are incorrect, it does not corrode Sn , steel, Fe or Pt .

A. L. HENNE

The decomposition of hypobromous acid. FRIEDRICH POLLAK AND ETELEA DOKTOR. *Z anorg allgem Chem* 196, 89-112 (1931)— HBrO is best prepd by treating concd AgNO_3 soln with pure Br_2 , followed by vacuum distn. A concn greater than 0.3 M will be unattainable because of the decompn of HBrO . Preps of HBrO always contain its decompn products, HBrO_2 and Br_2 , which fact has not always been considered in previous work. If HBrO is prepd from AgNO_3 , it contains HNO_3 which comes over during the distn. This HNO_3 can be detd analytically only by the nitrometer method. The nitrometer method of Klemenc and Hayek (*C. A.* 21, 3285), which is especially suitable for a small quantity, was modified for use in the presence of the Br compd present in the prepn. Inexplicable irregularities in the distn and the decompn of HBrO make it almost impossible to prep a soln of definite concn. This decompn is greatest during the first part of the distn. The decompn of HBrO proceeds by 2 simultaneous, independent reactions. (I) $5\text{HBrO} \rightarrow \text{HBrO}_2 + 2\text{H}_2\text{O} + 2\text{Br}_2$ and (II) $4\text{HBrO} \rightarrow \text{O}_2 + 2\text{H}_2\text{O} + 2\text{Br}_2$. In the dark, at 15°, reaction (I) predominates. H^+ ions were not observed to influence it. Bubbling of N_2 through the soln did not affect it. HBrO_2 may be involved transiently in the decompn, but it did not appear in measurable concn.

R. H. LOMBARD

Composition of the deposit forming on zinc immersed in copper sulfate solution. I. A. GALECKI AND J. TOMASZEWSKI. *Roczniki Chem* 10, 437-68 (in German 468-71) (1930), cf. following abstr.—The coherence of the deposit forming on zinc rods immersed in CuSO_4 soln increases with stirring and with the concn and decreases with the acidity of the solns; at the same time the color of the deposit becomes lighter. The deposit contains two constituents, (a) a metallic, flexible, red, yellow or white coating, possessing either a dull or a polished surface, and (b) a dark brown to black powdery or flocculent ppt. The former consists of metallic Cu and Zn in various proportions, probably both as a mixt. and as alloys, while the latter contains oxides of the two metals, as well as basic sulfates and the free metals.

B. C. A.

Composition of the deposit forming on zinc immersed in copper sulfate solution. II.

A. GALECKI AND J. TOMASZEWSKI *Roczniki Chem.* 10, 601-27 (in German) (28.9) (1930), cf. preceding abstr. — With the increase of temp. the content of Cu in the ppt. decreases and the content of Zn increases. The neg. balance of the reaction is greater the higher the temp. Temps. at which the expt. was performed (15-50°) had no noticeable influence on the appearance and color of the ppt. The addn. of ZnSO₄ to the reaction system lowers the percentage of Zn in the ppt. by an amt. proportional to a certain extent to the amt. of ZnSO₄ added. The neg. balance of the reaction increases with added ZnSO₄. In a quiet electrolyte the addn. of ZnSO₄ has in general no noticeable influence on the external form and consistency of the ppt. in a stirred soln. addn. of ZnSO₄ causes the formation of more fragile ppts. A considerable portion of Zn is most probably in the form of basic compds. which dissolve in NaOH. In general, the ppt. varies greatly in form and compn. with varying conditions. J. KECRAA

The non-existence of the suboxides of bismuth, BiI₂ and BiI. V. CAGLIOTTI *Gazz. chim. ital.* 60, 443-5 (1930). For a long time there has been evidence both for and against the existence of Bi suboxides lower than BiI₂ (cf. *Pogg. Ann.* 107, 601 (1859), Guthmann and Herz *Z. anal. Chem.* 36, 422 (1908), Marino and Becarelli, *C. A.* 7, 1457) and it was therefore of interest to exam. BiI₂ and the reputed compds., BiI₂ and BiI, by x rays. BiI₂ and BiI were prepd. with the app. used by Marino and Becarelli (see cit.) for studying the system Bi-I. The powder method was used in the x ray examn., and the diffractograms are reproduced. No evidence of the existence of any compd. other than BiI₂ was obtained. A mixt. of BiI₂ and Bi in proportions corresponding to BiI₂ showed no interference lines of Bi, so that the substance which was supposed to be BiI₂ is probably a mixt. of BiI₂ and Bi. A mixt. of BiI₂ and Bi in proportions corresponding to BiI showed the lines of BiI₂ and Bi. This obscuring of the lines of one substance in a mixt. is being studied further, and it has already been found that Cu contg. 7% of Co-Si does not show the interference lines of the Co-Si. C. C. DAVIS

Silicic acids. H. W. DELTHEY AND W. NAGEL *J. prakt. Chem.* 129, 178-88 (1931) cf. *C. A.* 23, 1583. — The earlier expts. were repeated with various solvents and at lower temps. The H₂O content of the silicic acids obtained depends on the temp. and is not much influenced by the proportion of the Ph₃COH in the reaction. At -18° to -20° and with 4.5 mols. of Ph₃COH a product with a H₂O content of 27% is obtained when using CHCl₃ as a solvent. This compares with the product obtained with Li₂O at 35°. The 27% is too high to assume the presence of the meta acid. With CCl₄ at -17° to -18° a product with 20% H₂O is obtained. This value is too low for the meta acid. No ortho acid could be obtained. The soly. of the silicic acids, detd. by dialysis, was not over 6% and decreased with age. The splitting off of H₂O is probably not intra but inter mol. F. SCHOTTE

The existence of violet tungsten oxide, W₂O₁₁. E. TARJAN *Naturwissenschaften* 19, 160-7 (1931). Expts. in H₂-H₂O mixts. between 700° and 1000° on W oxides did not confirm the existence of W₂O₁₁ (blue), in agreement with the results of Reinders and Verloet (*C. A.* 18, 349). Inside the range of existence of this compd. (Chaudron, *C. A.* 14, 2451) a violet product was obtained of compn. W₂O₁₁ (oxidation value 1.70 to 1.84%), i. e. between W₂O₇ and W₂O₉. Above W₂O₁₁ blue reduction products are formed at low H₂ concns. with oxidation values of a few tenths of a percent; below W₂O₁₁ no blue products could be obtained, further reduction gave W₂O₇ and W₂O₉ can be prepd. from W₂O₇ as well as from W₂O₉ at 1° to 30% H₂ at 900°. In CO-CO₂ mixts. W₂O₁₁ can be made inside the alleged W₂O₉ limits from either side (900°, 25% CO₂) (cf. van Liempt, *C. A.* 16, 1530). W₂O₁₁ crystallizes in red violet needles with a metallic luster; it has a triclinic lattice different from that of W₂O₇ and W₂O₉. Ignition of a pressed mixt. of 3W₂O₇ + 1W₂O₉ in an inert atm. gives the violet oxide in a manner similar to the prepn. method of W bronzes. The relationship with these bronzes is furthermore evident from the chem. reactivity and the metallic cond. ($\rho = 2 \cdot 10^{-4}$ ohm cm for a pressed sample, sintered at 1200°). Leiser (*C. A.* 2, 760) has described the electrolytic prepn. of W₂O₁₁; earlier erroneous data on the W oxides are partly due to retarded reduction processes and hence apparent equilibria. B. J. C. VAN B.

The preparation of complex bromine derivatives of quaternary tungsten. H. PAULSEN & BECK *Z. anorg. allgem. Chem.* 196, 55-8 (1931), cf. Collenberg *C. A.* 12, 2501. — The following complex bromotungstites were prepd. by methods which in principle consisted in treating the complex K, NH₄, or Na oxalotungstites with concd. HBr, sepg. excess oxalate and bromide and treating with a soln. of the bromide of the metal or base in HBr. *Dirubidium pentabromotungstite*, Rb (WOBr₅), bright, olive-brown crystals; *dicerium pentabromotungstite*, Cs (WOBr₅), greenish yellow powder of microscopic octahedra; *monopyridine tetrabromotungstite* (C₅H₅NH) (WOBr₄), brownish green

powder of square platelets *diammonium pentabromotungstate*, $(\text{NH}_4)_2(\text{WOBr}_5)$, olive-green octahedra, decompd on drying, *monotetraethylammonium aquotetabromotungstate*, $\{(\text{C}_2\text{H}_5)_4\text{N}\}(\text{WOBr}_4 \cdot \text{H}_2\text{O})$, bright green powder of cryst scales. The 3 series of salts, $\text{Me}_2(\text{WOBr}_5)$, $\text{Me}^1(\text{WOBr}_4 \cdot \text{H}_2\text{O})$ and $\text{Me}^1(\text{WOBr}_5)$, are analogous to the previously known chlorotungstites and have the same cryst form as the latter, although their color is more toward brownish yellow. The bromotungstites are less sol in concd HBr than are the corresponding chlorotungstites in concd HCl. The color of the bromotungstite soln is green, that of the chlorotungstite soln is blue. Treatment of Na oxalotungstite with concd HIF gives a violet-colored soln. The bromotungstites are more susceptible to hydrolytic decompn than are the chlorotungstites, but they may be preserved in dry CO_2 . R. H. LOMBARD

Sulfur derivatives of perhenic acid. WILH. FEIT *Z. anorg. Chem.* 44, 65-6 (1931).— H_2S is passed into a satd soln of KReO_4 . The soln becomes yellow in color at first, but after a day's flow of gas becomes deep brown, almost opaque. When no further change is observed the H_2S is removed by boiling, and the soln is carefully concd by evapn at 35° . When only about $1/4$ the original volume remains, small brown crystals sep which appear to be KReO_4 contg a small quantity of this salt. On further concn the crystals sepg become even darker and are accompanied also by a dark brown ppt. which is probably a sulfide of Re. The soln is finally dried without further sepn of crystals to a dark brown mass which is easily sol in H_2O and contains 14-15% S. It is a mixt of KReO_4S , KReO_4S_2 , KReOS_2 , KReS_4 . A soln of this residue is treated with N/20 TiSO_4 and the dark brown ppt. is filtered, washed and dried at 35° . Analysis gave the compn TiReS_4 . The filtrate from this still contained Re and S, and on further treatment with TiSO_4 yielded a bright yellow cryst. ppt of TiReO_4S . A soln of this salt on treatment with HCl gave a fugitive red color followed by sepn. of S, showing that free HReO_4S is unstable. H. STOERTZ

The hypophosphates of cesium and rubidium. L. HACKSPILL and J. WEISS *Compt. rend.* 192, 425-6 (1931).—The hypophosphates of Cs and Rb (POH_2M) are similar to those of the other alkalis. Their density and their affinity for H_2O increase with their at. wt. and their decompn with heat is similar. AMY LEVESCONTE

Study of the absorption spectra of various series of rare earth double nitrates. DENIS W. PEARCE and J. ALLEN HARRIS *Trans. Roy. Soc. Can.* [3], 24, Sect. 3, 145-51 (1930).—A study was made of the various double nitrates of the rare earth elements to ascertain if it was possible that some might be more efficient than others in effecting a sepn of these elements, particularly of II from Nd and Sa. From a study of the relative intensities of the most characteristic absorption band of each element it would appear that (1) The double NH_4 nitrates afford the most efficient sepn of La and Pr from Nd. (2) The double Mg and double Co nitrates are of the same efficiency in their relative sepn's, no definite sepn being possible by these salts. The ease of crystn renders them useful for preliminary sepn of the more basic members. (3) The double Mn nitrates are to be recommended as a means for sepg Sa and Nd. (4) The double Cu nitrates appear to effect a reversal of soly from the usual serial order of the double nitrates, ϵ , according to at. wts. (5) Pb appears to produce little effect on the sepn of the more basic members of the Ce group. Indications are afforded that $\text{Pb}(\text{NO}_3)_2$ may be helpful in the sepn of II. J. W. SHIPLEY

Observations on the complex cyanides of nickel, copper and chromium. W. GLUUD and W. RIESE *Ber. ges. Kohlentechnik* 3, 452-9 (1931).—In strongly ammoniacal solns, $\text{K}_2\text{Fe}(\text{CN})_6$ decomposes as follows under the influence of AgNO_3 . First one mol of NH_4 is taken up $\text{K}_2\text{Fe}(\text{CN})_6 + \text{NH}_3 \rightarrow \text{KCN} + \text{K}_4[\text{Fe}(\text{CN})_6 \cdot \text{NH}_3]$. The AgNO_3 removes the cyanide and takes the reaction to the right in steps as follows: $\text{K}_4[\text{Fe}(\text{CN})_6 \cdot \text{NH}_3] + \text{NH}_3 \rightarrow \text{KCN} + \text{K}_4[\text{Fe}(\text{CN})_6 \cdot (\text{NH}_3)_1]$, $\text{K}_4[\text{Fe}(\text{CN})_6 \cdot (\text{NH}_3)_1] + [\text{Fe}(\text{CN})_6 \cdot (\text{NH}_3)_1]$ and finally $\text{Fe}(\text{CN})_6$ are formed. Similar reactions occur with Ni salts, $\text{K}_2[\text{Ni}(\text{CN})_6]$ being converted into $\text{K}[\text{Ni}(\text{CN})_6 \cdot \text{NH}_3]$, $[\text{Ni}(\text{CN})_6 \cdot (\text{NH}_3)_1]$ and finally also into $[\text{Ni}(\text{NH}_3)_2\text{CN}](\text{CN})$ and $[\text{Ni}(\text{NH}_3)_4](\text{CN})_2$. Similar results are obtained with Cu and Cr compds. With $(\text{NH}_4)_2\text{S}_x$ these complex cyanides lead to the formation of thiocyanates, the Ni salts furnishing typical reactions. If the formation of $[\text{Ni}(\text{NH}_3)_4](\text{CN})_2$ is assumed, this reacts with $(\text{NH}_4)_2\text{S}_x$ to form NiS and NH_4CNS . The corresponding Cu and Cr salts are formed somewhat less readily. The presence of concd NH_4OH suppresses any considerable thiocyanate formation. Varied reactions are given with FeSO_4 , the Ni salts giving a ppt. of $\text{Fe}[\text{Ni}(\text{CN})_6]$, no Fe whatever replacing Ni in the anion. With Cu salts, however, ferrocyanides are produced and also the following compd., $\text{KFe}[\text{Cu}(\text{CN})_6]$. With the complex Cr cyanide the reaction is as follows in alkaline soln: $\text{K}_2[\text{Cr}(\text{CN})_6] + 3\text{KOH} \rightarrow \text{Cr}(\text{OH})_3 + 6\text{KCN}$. The excess of KCN then forms a ferrocyanide with the ferrous salt. H. STOERTZ

Phase rule studies on metallic thiocyanates (OCCLESIAHAW) 2.

LÖWENSTEIN, ELSE Über Subhaloide der Homologen des Quecksilbers. Freiburg: B. Speyer und Kaerner 32 pp M 150

7—ANALYTICAL CHEMISTRY

W. T. HALL

Quantitative emission spectrum analysis in any percentage without calibration curve. G. SCHRIEDT and O. SCHNETTLER *Naturwissenschaften* 19, 134 (1931).—Certain phases of a method previously described (*C. A.* 24, 5622) have been checked by the use of a Pt absorption wedge. The results are briefly discussed. B. J. C. v. d. H.

Physical methods in the chemical laboratory. XVII. A method for increasing the accuracy in quantitative emission spectral analysis and its testing. G. SCHRIEDT, C. F. LINSTRÖM and O. SCHNETTLER *Z. angew. Chem.* 44, 145-51 (1931), cf. *C. A.* 25, 1440.—The photographic plate, with respect to its applicability for spectrographic analyses, was studied, with particular attention to the dependence upon the wave lengths. With the aid of a thermoelec. photometer, which is suitable for technical purposes, the behavior of the intensity of the lines was carefully studied, and as a result of this study a procedure was developed which will serve for the spectrographic detn. of varying quantities of substances with an accuracy of about 3%. W. T. H.

Colorimeters, spectrophotometers and nephelometers. C. DIGAUD *Ann. chim. anal. chim. appl.* 13, 1-5, 33-54, 65-72 (1931).—This is an account of the marked advances that have been made in the field of analytical chemistry, particularly with respect to applying the principles of modern phys. chemistry. A discussion of the theory underlying all processes ending with visual observation begins with an explanation of the anatomy of the eye, then turns to photoelec. cells, thermoelec. couples, thermoelec. piles and photographic plates and finally compares the sensitiveness of the eye with that of the photoelec. cell and of the photographic plate. The app. and its use are described in detail. W. T. H.

Nephelometric titrations. II. Standard-solution end-point. CLYDE R. JOHNSON *J. Phys. Chem.* 35, 530-5 (1931).—It has been pointed out (*C. A.* 25, 1178) that there is a possible error in the equal-opalescence end point in nephelometric detns. To avoid this possible error it is proposed to work to a different end point called the "standard soln. end point." For this 2 sets of standard solns. are necessary. In the titration of Cl^- with AgNO_3 , e. g., the necessary quantity of AgNO_3 , to within a few tenths of a mg., is added, and then both Cl^- and Ag are detd. nephelometrically in the supernatant soln. by use of solns. satd. with AgCl ppt. as standards. W. T. H.

The stability of standard arsenious acid solutions. IV. TANANAEV. *Ukrainskii Khim. Zhurnal* 5, Sci. Pt., 217-26 (1930).—(1) Standardized arsenious acid loses the strength of its titer through oxidation to arsenic acid. The oxidation rate is directly proportional to the alk. of the soln. A 0.1 N soln. of As_2O_3 in 1 N NaOH suffers a daily loss in titer equaling approx. 0.176% of As_2O_3 . (2) Weakly alk. solns. ($\text{pH} = 7-9$) remain unchanged for 15 years. (3) NaHCO_3 free from Na_2CO_3 is most suitable and can be prepd. by passing CO_2 through the soln. (4) Solns. of As_2O_3 prepd. from the pure chem. remain sterile and do not form As_2S_3 . B. S. LEVINE

Systematic qualitative analysis with small quantities of cations. A. SCHEINKMANN *Z. anal. Chem.* 83, 176-88 (1931).—Instead of starting with 1 g. or more of material, only 0.1-0.15 g. of sample is taken, and the reagents are added in drops, an unnecessary excess being avoided. The first group of AgCl , PbCl_2 and Hg_2Cl_2 is treated with a little 5% KNO_3 soln., in which PbCl_2 is more sol. than in water, but is otherwise analyzed in the conventional manner. For pptg. the second group, a satd. soln. of H_2S is used, which is prepd. daily but which is much more convenient to use than the gas itself. The sulfide ppt. is treated with KOH soln. to dissolve the sulfides of As, Sb, Sn and Hg. From this soln. HgS is pptd. by NH_4Cl , and the sulfides of As, Sb and Sn by AcOH , after which this subgroup is analyzed conventionally. The sulfides of Group II which are insol. in KOH soln. are dissolved in HCl and H_2O_2 and treated with NH_4OH , after which the procedure is normal. In Group III, the NiS and CoS left undissolved by dil. HCl are dissolved by treatment with AcOH and H_2O_2 , which is decidedly advantageous. In testing for Al, both the morin and alizarin tests are used. For Zn the resorcinol and ferrocyanide tests are used, and for Mn the green fusion test and the treatment with PbO_2 and HNO_3 are prescribed. The filtrate from this

Fe-Al group is evapd. to dryness after making acid with HCl and ignited with oxalic acid, whereby insol. BaCO_3 , SrCO_3 , CaCO_3 and MgO are formed together with water-sol. alkali carbonates. In this way the baneful effects of excess NH_3 , salts are overcome. These are the principal ways in which the procedure differs from that ordinarily used. Several lab. devices are shown which will be found useful. W. T. H.

Mercurimetric studies. O. TOMÍŠEK AND O. PROČKE *Collection Czechoslov. Chem. Comm.* 3, 116-25 (1931).—Votoček found that $\text{Na}_2\text{Fe}(\text{CN})_6(\text{NO}) \cdot 2\text{H}_2\text{O}$ could be used as indicator in titrating Cl^- and Br^- but not I^- with $\text{Hg}(\text{NO}_3)_2$. With respect to the solv. of the Hg nitroprussiate and the extent of its ionization, but little data are available. It would seem probable that the molar solv. is not greater than 2×10^{-4} and that its soly. product is below 10^{-8} . The physicochem. studies here described lead to the same conclusion that the molar solv. is below 5×10^{-4} . Other observations made will be given a practical application in a subsequent paper. W. T. H.

Dimethyl-p-phenylenediamine hydrochloride for the determination of small quantities of chlorine. KNUT ALFTHAN *Finska Kemistisamfundets Medd.* 36, 109-12 (1927).—The reagent gives with Cl the same shade of red color as methyl red (dimethyliminobenzene-o-carboxylic acid) shows in acid soln. The water to be tested is treated with the reagent and the resulting color is compared with those of known solns. of methyl red. In testing drinking water 0.01 mg. Cl can be detd. J. RYSELIN.

Indirect estimation of silicon in 45 to 52 percent ferrosilicons. G. T. DOUGHERTY *Ind. Eng. Chem., Anal. Ed.* 3, 138-9 (1931).—The direct detn. of Si in Fe-Si requires 15-2 days. To avoid car demurrage, an indirect method has been worked out requiring but 0.5 day and which depends upon carefully treating the sample with HF and H_2SO_4 in a Pt or palau crucible and weighing the residue after evapn. and ignition. W. T. H.

Lithium chloroplatinate and the separation of potassium from sodium and lithium by the unmodified, original, Fresenius method. G. FREDERICK SMITH AND A. C. SUEB *J. Am. Chem. Soc.* 53, 947-57 (1931).— $\text{Li}_2\text{PtCl}_6 \cdot 6\text{H}_2\text{O}$, prepd. from Li_2CO_3 and H_2PtCl_6 , and twice recrystd., can be used to advantage for pptg. K_2PtCl_6 . The ppt. then has the normal concn., which is not true when H_2PtCl_6 is used as precipitant. Moreover, the ppt. is stable at 200°C , whereas the ppt. produced with H_2PtCl_6 decomposes above 100° . For washing the K_2PtCl_6 ppt., 80% LiOH is recommended. The errors due to the use of H_2PtCl_6 are due to decompn. of the reagent, which does not take place when recrystd. Li_2PtCl_6 is used. W. T. H.

Quantitative determination of calcium, magnesium and phosphorus in feedstuffs and cattle excreta. Improved technic. H. P. MORRIS, J. W. NELSON AND L. S. PALMER *Ind. Eng. Chem., Anal. Ed.* 3, 104-7 (1931).—Detailed directions are given for fusing the ash with soda, extg. the melt with HCl and water and detg. Ca, Mg and P in the soln. To avoid the necessity of removing all SiO_2 , the CaCO_3 ppt. is titrated rather than weighed. The value of the method is shown by exptl. data. W. T. H.

The selenious acid method for the determination of zirconium. STEPHEN G. SIMPSON AND WALTER C. SCHUMB *J. Am. Chem. Soc.* 53, 921-33 (1931).—For decompg. Zr minerals, the Muehlberg method of fusing with Na_2O_2 and C is recommended, the fusion being made a second time with any residue which is insol. in acid. After the fusion, ext. the melt with water and reject the soln. of sol. Na salts. Dissolve the residue in HCl, filter if necessary and evap. the soln. to fumes with 25 cc. of 18 N H_2SO_4 . Cool, dil. and filter off SiO_2 . Make the filtrate ammoniacal, filter and wash slightly. Promptly dissolve the ppt., which contains ZrO_2 , in 15 cc. of 12 N HCl, add 20 cc. of 10% LiOH , heat nearly to boiling, dil. to 500 cc., again boil and add 20 cc. of 10% H_2SeO_4 soln. When the liquid is clear, filter off the ppt. of $4\text{ZrO}_3 \cdot 3\text{SeO}_3 \cdot 18\text{H}_2\text{O}$, and after very slight washing, dissolve it in 15 cc. of 12 N HCl. Add 20 cc. of 3% H_2O_2 to reduce any quadrivalent Ce, dil. to 500 cc. and again ppt. with H_2SeO_4 . Digest the 2 filters used for the Zr ppts. with 40 cc. of hot 10% oxalic acid and add the soln. to the main ppt. contg. Zr. Dil. to 200 cc., heat to boiling and add 12 cc. of 6 N HCl. After it has stood at least 10 hrs., filter and wash the ppt. with 2.5% oxalic acid dissolved in 0.24 N HCl. To the filtrate contg. the Zr add 30 cc. of 18 N H_2SO_4 and evap. till all the $\text{C}_2\text{O}_4^{2-}$ has been destroyed, as shown by no more evolution of CO and CO_2 . Filter off any Se ppt. and make the filtrate ammoniacal. Filter off the $\text{Zr}(\text{OH})_2$ ppt., and after a little washing dissolve it in 15 cc. of 12 N HCl, add 20 cc. of 3% H_2O_2 , dil. to 500 cc., add 20 cc. of 10% H_2SeO_4 soln., filter, ignite the basic Zr selenite over the blast lamp and weigh as ZrO_2 . W. T. H.

Separation of calcium and magnesium by the molybdate method. R. C. WILEY

Ind. Eng. Chem., Anal. Ed. 3, 127-9 (1931).— CaMoO_4 is insol. in water, whereas MgMoO_4 is quite sol. Quantities of Ca ranging from 20 to 50 mg. can be sepd. from 12-70 mg. of Mg by treating the soln. (0.06 *N* in Ca, 0.075 *N* in Mg, or less) at the boiling temp. with a slightly ammoniacal soln. of NH_4 molybdate (about 0.4 *N*), added dropwise. The pptn. of the Ca is complete if boiling is continued 1.5 hr. the supernatant liquid is clear. The ppt. can be weighed as CaMoO_4 after drying at 130° for 30 min. The Mg can be detd. as MgNH_4PO_4 in the filtrate. W. T. H.

Potentiometric determination of iridium. SPENCER WOO AND DON M. YOST *J. Am. Chem. Soc.* 53, 854-8 (1931), cf. C. A. 25, 1725.—The iodometric method, in which quadrivalent Ir is reduced to the trivalent Ir by means of I^- in 0.1-0.2 *N* HCl and the liberated I_2 is titrated with $\text{Na}_2\text{S}_2\text{O}_3$, gives good results if benzene is used to show the end point. Good results are also obtained by direct titration with TiCl_3 . The potentiometric end point is clear in 0.1-0.2 *N* HCl if 2-3 g. of NaCl is added and the TiCl_3 is added slowly toward the last. Directions are given for prep. chloro-iride and chloro-iridium acids. W. T. H.

Method for the separation and gravimetric determination of osmium. RALPH G. GUNNERY *Bur. Standards J. Research* 6, 421-49 (1931).— Os , when present as chloride or bromide, can be pptd. completely as hydrated OsO_4 from boiling solns. of pH 1.5-6.3. The presence of alkali chloride or sulfate does no harm. The marked tendency to spatter when the crude is heated can be overcome by impregnating the ppt. with NH_4Cl . Os , when reduced in H_2 and cooled to room temp. in CO_2 , remains unattacked by air for a long time. Os , when present as bromosulfate or alkali osmate, can be volatilized completely as OsO_4 by boiling in 1% HNO_3 . When present as chlorosulfate, it cannot be eliminated easily in this way, but from concd. H_2SO_4 , near its b. p., the OsO_4 can be distd. off from the latter concd., but less readily from the former concd. By a mixt. of H_2SO_4 and HNO_3 it is possible to effect complete volatilization from either type of concd. No RuO_4 is volatilized from 40% HNO_3 or from concd. H_2SO_4 alone, but from concd. H_2SO_4 contg. HNO_3 or from concd. HNO_3 alone RuO_4 is gradually volatilized. A soln. of 6 *N* HCl satd. with SO_2 is a suitable reagent for recovering the distd. OsO_4 . It is necessary, however, to evaporate repeatedly with concd. HCl to decompose the Os sulfate concd. completely. The thioarsa test for the detection of Os can be made sensitive to 0.52 mg. of Os in 1.0 ml. of soln. In the app. recommended for the detn. of Os , all connections are of ground glass. The 700 ml. distg. flask carries a tap funnel in the neck, a glass tube extending to the bottom of the flask for the entrance of the gas, and a gas exit tube leading to a second flask, contg. 170 ml. of 6 *N* HCl which has been satd. with SO_2 . In this first receiver practically all of the OsO_4 is condensed, but for the sake of safety there are 2 other flasks provided, each contg. 50 ml. of the SO_2 -HCl soln. Between the first and second flasks, a small tap funnel is fused into the gas line, through which added HClSO_3 can be introduced if necessary. It was found necessary to keep the stopcock of this funnel greased, but otherwise no lubricant is advisable. To the soln. in the distg. flask, add, if necessary with 100 ml. of water, add 40 ml. of 7.5 *N* HNO_3 and flush out the funnel with 10 ml. of water. Pass a slow current of air through the app. and heat the soln. to boiling. Distil for 1 hr. This should be sufficient to remove Os from alkali osmate or bromosulfate, but 7-8 hrs. are necessary if chlorosulfate is present. In this case it is better to distil from concd. H_2SO_4 or if Pu is absent, from concd. H_2SO_4 to which a little concd. HNO_3 has been added. Unite the several portions of HClSO_3 and evaporate on the water bath in an unetched beaker. Digest the residue with 10 ml. of concd. HCl for 15 min. and evaporate a second time. Repeat the digestion and evapn. 3 times more. Take up the final residue in 170 ml. of water and heat to boiling. Add NaHCO_3 soln. until a ppt. appears and suddenly coagulates. Add bromophenol blue indicator and enough more NaHCO_3 soln. to impart a faint, bluish color to the soln. Boil 5-6 min., filter through a Munnor crucible and wash the ppt. with hot 1% NH_4Cl . Cover the washed ppt. with cold NH_4Cl , add a few drops of the wash liquid, and apply suction to get the NH_4Cl into the ppt. Wipe off the coating of NH_4Cl that comes through the crucible, cover with a Rose crucible lid of quartz and heat in H_2 . At the end heat strongly for 10 min. and allow to cool in H_2 , finally replacing with CO_2 . Weigh as crystalline Os . W. T. H.

New method for reduction of tin and antimony prior to titration. B. S. EVANS, *Analyst* 56, 171-7 (1931).—A soln. of 5 g. $\text{NaH}_2\text{PO}_4 \cdot \text{H}_2\text{O}$ in 6 *N* HCl, added to about 100 ml. of 6 *N* HCl contg. 0.002-0.005 g. of Sn and 1 cc. of satd. H_2Cl_2 as catalyst, will reduce all Sn^{++++} to Sn^{++} if boiled 15 min. in an atm. of CO_2 . Then, by dilg. with 25 times as much of a boiled and cooled soln. of 10 g. citric acid, 0.4 g. KI and some starch, the Sn can be detd. by titration with standard I_2 . Particular pains were taken to avoid all contact with air. The reduction of Sb to the trivalent condition takes

place with 5 min. boiling and without the use of any catalyst. Fe, Cu and As must be removed. The results obtained were excellent. W. T. H.

Determination of lead as chromate in presence of perchloric acid. D. J. BROWN, JOSEPH A. MOSS AND JOHN B. WILLIAMS. *Ind. Eng. Chem., Anal. Ed.* 3, 134-5 (1931) — $PbCrO_4$ is less sol. in dil. $HClO_4$ solns. than it is in dil. HCl or dil. HNO_3 . Directions are given for carrying out the gravimetric and volumetric detn. of Pb as $PbCrO_4$, after first removing HCl and HNO_3 by evapn. with $HClO_4$, and 24 results obtained with a 0.5 g. sample of a certain ore contg. 82.04% Pb showed Si 55-82.04%. W. T. H.

Photoelectric microanalysis. MOTOTARŌ MATSUI AND TŌKITI NODA. *J. Soc. Chem. Ind. Japan* 33, Suppl. binding 517-8 (1930) — Cu in Cu-NH₃ soln. can be detd. by measuring with a photoelec. cell the amt. of light transmitted through such a soln. The light source was an incandescent lamp. The thickness of the Cu-NH₃ soln. was 12 cm. If I is the current from the photo-elec. cell when the Cu-NH₃ soln. is in the light path, and I_0 the current when H_2O is substituted for Cu-NH₃ soln., a quant. relation exists between $I_0 - I$ or $\log(I_0/I)$. The photoelec. current was reduced 73.5% when a 0.00063 N Cu-NH₃ soln. was substituted in the cell by H_2O . An accuracy of about 5% was possible. The photoelec. cell used was of the A-filled Cs type. The tension was 80 v. in series with a 5 megohm protector. E. M. SYMONS.

Detection of cadmium. J. STANTON PIERCE AND W. T. FORSEE. *Ind. Eng. Chem., Anal. Ed.* 3, 188-9 (1931) — Instead of using KCN or Fe powder before testing for Cd in the usual qual. scheme, it is recommended to boil the neutral or slightly acid soln. for 3-5 min. with 1-5 g. of fine Ni powder which serves to reduce the Cu. Then, after filtering, a soln. is obtained which can be tested for Cd by adding $AcOH$ and H_2S . W. T. H.

Detection of zinc and cadmium by the cyanide method. N. A. TANANAYEV AND N. S. FEDULOV. *Ukrainski Khim. Zhurnal* 5, Sci. Pt., 213-6 (1930) — To a soln. of the substance add a considerable excess of NH_4OH , heat and filter. If Mn is present add H_2O_2 and filter. To the filtrate add a soln. of KCN until the colors due to Ni, Co and Cu disappear. Add Na_2S and heat. Wash the ppt. 3 or 4 times with water. A white ppt. which completely dissolves in an excess of KCN indicates the presence of ZnS_2 alone. If the ppt. is yellow, CdS is present alone or in conjunction with ZnS . To establish this fact add an excess of KCN and acidify the soln. with $AcOH$. A white ppt. or turbidity indicates the presence of ZnS . There is some incomplete exptl. evidence to show that the method works in the presence of Pb and Sb when the NH_4OH added is not too strong. B. S. LEVINE.

Qualitative tests for zinc, copper and silver. A. SERGEEV. *Ukrainski Khim. Zhurnal* 5, Sci. Pt., 227-9 (German abstract 230) (1930) — Add a few drops of concd. $NaOH$ and some pulverized Al to the soln., and stir with a glass rod. A change in color and a swelling of the Al particles indicate the presence of Zn. To a neutral or slightly ammoniated soln. of the substance add some Hg . Formation of gray or silvery needle-like crystals indicates the presence of Ag. Heat the end of a narrow glass test tube and apply to a drop of the soln. to be tested. Heat the end of the tube for about 2 min. in the lower part of a Bunsen flame. A copper-orange filament is formed if Cu is present. B. S. LEVINE.

Pyrognostic assaying of platinum and palladium. HERCULANO CALMON, *Rev. bras. chim.* 1, 84-5 (1929) — C. describes his method for the detection of Pd and Pt, which is based on the incandescence that is observed in the ashes of the filter paper or in asbestos threads that have been impregnated with the salt to be investigated, heated to a red heat and put in the flame of illuminating gas, alc. vapors or a mixt. of $EtOH$ and sulfuric ether. Working out a metallic web is recommended. The reaction is very sensitive. JOHN M. LADINO.

New method for the titration of iron according to J. Knopp. OTTO ROTH AND AGOSTO PRO SOBARINO. *Rev. bras. chim.* 1, 129-49 (1929) — For detg. Fe^{++} , the following conclusions are drawn. The use of diphenylamine indicator with dichromate is superior to the use of permanganate. The metals of the second group must be removed with H_2S , especially As and Sb. Min. quantities of Cu cause no error. Colloidal silica or other metals such as Hg and Sa need not be removed. Colloidal Pt caused by fusion of the oxide in a Pt crucible is harmful, making the use of permanganate necessary. A Ag crucible is to be preferred. Care must be taken not to reduce any $AgCl$, the reduction of which takes place quite readily. JOHN M. LADINO.

Determination of aluminum oxide in aluminum and its alloys. A. M. SHANDOROV. *Tretnise Metallur.* 1930, 672-9, *Chimie & industrie* 25, 37 (1931) — S. verified the accuracy of the Cl_2 gas method, using a gas heated, transparent quartz tube so as to follow the reaction, which is strongly exothermic and requires careful regulation of the

heating and Cl feed. It is essential that the Cl be free from O_2 and water. After the chlorination the boat is ignited at 1000° to burn the C. The analysis requires about 1 hr. S found 0.07-0.10% Al_2O_3 in com. Al and 0.07-0.21% in duralumin.

A. PAPINEAU COUTURE

Analysis of metal containing copper. I. Analysis of copper. II. ARAI, C. UYET-SUKI AND S. TACHIBANA. *Repts. Imp. Ind. Research Inst., Osaka, Japan* 11, No. 15 (1931). The procedure for the electrolytic detn. of Cu is given. F. I. NAKAMURA

Supplement to the new iodometric method for determining vanadium in alloy steels and in ferro-vanadium. W. WEAR. *Z. anal. Chem.* 83, 161-4(1931); cf. C. A. 25, 1. For the decompn. of excess $(NH_4)_2S_2O_8$, 15 min. is inadequate, but 30-35 min. should suffice. The reaction between Ie^{+++} and I is so slow in the presence of H_3PO_4 , that as much as 3 g. Ie liberates I_2 corresponding to not more than 0.2 cc. of 0.05 N $Na_2S_2O_3$, when 0.2-0.3 g. of KI is added, with 1-2 g. of KI about twice as much I_2 is liberated. When oxalic acid is present, more I_2 is set free. With highly alloyed steels this effect is less, and by keeping the concn. of the H_3PO_4 soln. low, the interference of Mn is lessened and the addn. of oxalic acid is unnecessary. There is, on the other hand, no danger of any Cr being oxidized by the persulfate when Ag ions are absent. When the Cr is high in an alloy steel, the Cr carbide should not be removed until after the oxidation with HNO_3 . If this carbide is not removed, it will tend to reduce vanadate later in the analysis. The persulfate should not be added until the nitrous fumes have all been boiled off from the soln. The procedure as modified is as follows: Dissolve 1-3 g. of steel in a 500-cc. Erlenmeyer flask in 25 cc. of H_3PO_4 (d. 1.70) and 200 cc. of water (10-15 min.). Oxidize with concd. HNO_3 and add 5 cc. in excess. Boil a few min. to remove nitrous fumes. Filter off any undissolved carbide. Add 10 cc. of 5% $(NH_4)_2S_2O_8$ soln. and boil 35 min. Add 25 cc. more of H_3PO_4 , cool to room temp., add 2-10 cc. of 0.2 N KI soln., shake well and after 5 min. titrate with 0.05 N $Na_2S_2O_3$ in a vol. of about 200 cc.

W. T. H.

Use of microanalysis in the streak test (for precious-metal alloys). R. STREIBER AND H. HOLZER. *Mikrochemie* [N. S.] 2, 264-70(1930).—The streak test is made on a roughened depression in a microscope slide. When the alloy is chiefly Ag, the streak is dissolved in HNO_3 on the slide, and the Ag is pptd. as chloride and removed by centrifuging. The filtrate is evaporated to dryness, the residue dissolved in water, and 1 drop of K mercurithiocyanate soln. is added; green crystals indicate the presence of Cu and brown crystals the presence of Cd. For the examn. of a Au streak it is dissolved in aqua regia; the metals are pptd. as sulphides by passing H_2S through the soln. from a capillary tube, and the ppt. is collected in a microcentrifuge, washed, digested with $(NH_4)_2S$ to remove the Au, and dissolved in HNO_3 . Min. drops of the soln. are tested for Pb by addn. of $AcOK$ and KNO_3 (triple K Pb Cu nitrate pptd.) for Bi with K sulfate (double sulfate pptd.), for Pd with dimethylglyoxime in $AcOH$, for Ni with the same reagent and NH_4 , for Co, Zn and Fe with NH_4 mercurithiocyanate, and for Al with alizarinsulfonic acid. Pt streaks may be similarly tested for the presence of Cu, Pd and Ag; Au is detected by the purple of the Cassius test.

B. C. A.

Titrimetric determination of small quantities of ammonia, with particular attention to water analysis. S. KCHNEL HAGEN. *Z. anal. Chem.* 83, 164-75(1931).—An app. is shown which is suitable for the detn. of free NH_3 in water. One of the principal advantages lies in the avoidance of rubber stoppers and rubber connectors and in the use of a quartz condenser tube. The NH_3 is caught in a measured vol. of $1/100$ N HCl and the excess titrated with NaOH of the same concn. against an indicator of 0.01% soln. of methyl red and 0.04% soln. of bromothymol blue in 90% alc. For the detn. of albuminoid NH_3 , it is recommended to heat with H_2SO_4 , 0.1 g. graphite, 0.75 g. $HgSO_4$, 1.00 g. $CuSO_4 \cdot 5H_2O$ and some K_2SO_4 , and finally to add Na_2S and NaOH and distil. By the app. and method described, 0.1-0.3 mg. of NH_3 can be detd. with an accuracy of about 5%.

W. T. H.

Indirect determination of magnesium carbonate in the presence of calcium carbonate. G. A. PAVCHENKO. *Ukrainski Khim. Zhurnal* 5, Sci. Pt., 187-95(1930).—Dissolve a weighed sample in an excess of standard acid and titrate the excess. Now add an excess of standard oxalic acid and titrate the excess of the oxalic acid with $KMnO_4$.

B. S. LEVINE

The estimation of sodium carbonate in sodium bicarbonate. A. K. BAKO. *Ukrainski Khim. Zhurnal* 5, Sci. Pt., 197-206(1930).—When the % of soda in the bi carbonate exceeds 3-4%, it can be detd. by titration against phenolphthalein, if a pure soln. of $NaHCO_3$ and an equal aml. of the indicator are used as a control, or by the method of Simpson with a mixt. of 6 parts of bromophenol blue and 1 part of cresol red

as indicator. If the Na_2CO_3 content is below 3%, a colorimetric detn. is better.

B. S. LEVINE

The oxalate method for the analysis of potassium thiocyanate using the borax titer. N. A. TANANAEV AND N. A. LAZARKEVICH *Ukrainskii Khim. Zhurnal* 5, Sci. Pt., 209-12 (1930).—By means of the oxalate method, convert KCNS into K_2CO_3 and titrate the latter with acid soln. standardized against borax.

B. S. LEVINE

Determination of phosgene. J. C. OLSEN, GEORGE E. FERGUSON, VICTOR J. SABETTA AND LEOPOLD SCHEFLAN *Ind. Eng. Chem., Anal. Ed.* 3, 189-91 (1931).—Four methods for detg. COCl_2 have been studied: (1) absorption in alc. soln. of NaOH and subsequent titration of chloride or of excess NaOH ; (2) absorption in ammoniacal AgNO_3 soln. and weighing of the AgCl formed after neutralizing with HNO_3 , or titration of excess Ag ; (3) absorption in a satd. soln. of aniline and weighing of the resulting diphenylurea; (4) passing of the COCl_2 through a 2% soln. of NaI in acetone and titration of the liberated I_2 . Methods (3) and (4) were found reliable in studying decomposition products of CCl_4 mixts. and in samples of COCl_2 contg. other Cl compds.

W. T. H.

The separation of phosphoric acid as bismuth phosphate in quantitative analysis. A. KEŠANS *Acta Univ. Lathensis Kim. Fakult. Seriya 1*, No. 4-5, Fasc. 2 Burtneva (in German 121-6) (in Lettish 65-121 (1929)).—At room temp., a satd. soln. of BiPO_4 in 0.25 N HNO_3 contained 0.6 mg. of P_2O_5 in 100 cc. A 0.5 N HNO_3 soln. under the same conditions dissolves 2.4 mg. of P_2O_5 . In the presence of the halogen or sulfate ions, the soly. increases, because a portion of the BiPO_4 becomes basic BiOX or $\text{BiO}(\text{OH})$, SO_4 . This difficulty can be overcome by using solid BiONO_3 as the precipitant. The sepn. is so complete that only 0.1 mg. of P_2O_5 in 100 cc. of the soln. remains. The main pptn. is carried out in a 2 N or 5 N HNO_3 (either hot or cold) by adding the solid BiONO_3 and diluting to a strength of 0.5 N . The mixt. is warmed for 6-8 hrs. on a water bath and allowed to cool until the next day. This insures greater pptn. The cations Ca and Sr should not be present in amts. greater than 70 mg./100 cc. of diluted soln., and the other cations, Na , K , Mg , Ba , Zn , Mn , Co , Ni and Fe , in concn. greater than 100 mg./100 cc. of soln.

RUSSELL C. ERM

A new method for the optical determination of atmospheric ozone. ANDRÉ I. DUMINOWSKI *Compt. rend.* 191, 859-61 (1930).—By use of a galvanometer sensitive to 10^{-9} amp., with photographic registration, the distribution curve of solar energy was recorded between 4800 and 12 000 Å. U. on several days from August to December, 1929. In interpreting the results it is necessary to consider at some definite wave length the intensity, I , the mass of air traversed, M , and the atm. d., Δ ; $\text{colog } I = \text{colog } I_0 + m\Delta$. The ozone is detd. in 4 min., and 10-15 detns. should be made at one time, as atm. changes during a series of measurements were the principal source of error. The work was done at Montpelier. The value of ϵ varied from 0.20 to 0.36 cm.

GERALD M. PETTY

New method for determining carbon monoxide by combustion with oxygen in the presence of a new catalyst of two substances. ALBERT SCHMIDT *Z. anorg. Chem.* 44, 152-5 (1931).— I_2O_5 is commonly used for the oxidation of CO but is not very stable at the necessary temp.; it tends to retain the liberated I_2 , and but 1 mol. of I_2 is formed by the action of 5 CO . Combustion of CO to CO_2 , by passing the dry gas over CuO at 300-400°, is always more or less incomplete without a catalyst, but if 1-2 mm. quartz grains are moistened with $\text{Cu}(\text{NO}_3)_2$ and then ignited, a layer of fine CuO on SiO_2 is obtained which converts CO to CO_2 very efficiently. The SiO_2 tends to adsorb CO , the CuO adsorbs O_2 and the catalyst does not retain the CO_2 formed, which can be absorbed in $\text{Ba}(\text{OH})_2$ soln. An app. is shown which is suitable for detg. CO and in which the reaction tube is heated electrically. At the end of the process the excess $\text{Ba}(\text{OH})_2$ is titrated.

W. T. H.

Determination of hydrochloric acid in the presence of hydrobromic and hydriodic acids. G. G. LONGINESCU AND TH. I. PIRTEA *Bul. chim. soc. romând stinte* 31, No. 4/6, 77-87 (1930); cf. *C. A.* 25, 894.—The method described depends upon first pptg. with AgNO_3 in dil. HNO_3 and weighing the ppt. This gives the wt. of AgCl together with AgBr and AgI . In a second portion of the same size, the ppt. is kept as far as possible in the beaker and washed by decantation to remove all excess Ag . The ppt. is then heated with 50 cc. of 7-8% NH_3 soln. and treated with a soln. of KBr or HBr , whereby any chloride of Ag is converted into AgBr . This ppt. is also weighed. Finally, when all 3 halides are present, a similar treatment is given to the ppt. except that KI is added after the treatment with NH_3 soln. In this case, the total halide content is detd. as AgI . The purpose of the NH_3 treatment is to dissolve all the AgCl .

and part of the AgBr, whereby the KBr or KI soln. becomes more effective in accomplishing the desired conversion. W. T. H.

Microdetermination of the calcium ion. MORSEMAN *Bull. soc. chim. biol.* 12, 1014(1930) of C. A. 25, 1458.—Ca is pptd. as the salt $K_2CaNi(NO_3)_6$ from which N is reduced to NH_3 , and detd. by titration. The amt. of Ca can then be calcd. Fe, Mg, Al etc. cause less interference in this method than in the tungstate method. From 0.3 to 1.0 mg. Ca can be detd. C. G. KING

A new reagent for fluorides. CAMILLO PERTUSI *Atti III congresso naz. chim. pura applicata* 1930, 573 5.—The reagent consists of a soln. of benzidine acetate prepd. from 1.84 g. benzidine, a little concd. HCl and distd. H_2O to make up to 500 cc., and a 0.02 N Hg succinimide soln. mixed in equal parts. The reaction is quite evident with a drop of 0.05 N alkali fluoride or 0.0004 g. of HF . The compn. of the ppt. in a complex mol. of 2 mols. of HF , HCN , CaH_2 , NH_3 , HF and 1 mol. of HgF_2 . F. M. S.

Detection of bromides by the drop method. A. V. PAVLINOVA *Ukrainski Khim. Zhurnal* 5, Sci. Pt., 231(German abstract 232)(1930).—The method is based on the fact that pernitric acid formed by the interaction of H_2O_2 and HNO_3 , liberates Br from bromides. The Br gives a pink color with eosin or fluorescein. The reaction is sensitive to 1:12,500. Chlorides do not interfere with the reaction. Iodides do, but they can be removed with Na_2SO_3 . B. S. LEVINE

Conductometric titration of sulfate and barium. I. M. KOLTHOFF AND TOMRU KAMEDA *Ind. Eng. Chem., Anal. Ed.* 3, 129-33(1931).—In the conductometric titration of $ZnSO_4$ with $BaCl_2$, the end point is very poor unless about 30% of a/c. is added. In all cases, however, the titration curve bends before the equivalence point is reached. Similar results were obtained in titrating sulfates of K, Na and Li with $BaCl_2$, as well in the reverse titration of $BaCl_2$ with alkali sulfate solns. The end points are reproducible, however, to within about 1% of the Ba or SO_4 content, so that the conductometric method will be useful in the analysis of very dil. solns. The errors are smallest with Li_2SO_4 . The cond. of the soln. in the presence of $BaSO_4$ ppt. is greater than when the ppt. is removed, possibly because of the charge residing on the $BaSO_4$ in suspension and because of its influence on the cell const. W. T. H.

Determination of nitrite and sulfite in the presence of one another in salt mixtures and in meat products. E. SZABO *Z. Untersuch. Lebensmittel* 60, 389-95(1930).—The salt is dissolved in a little H_2O , Na_2CO_3 is added and the sulfite pptd. by a cold soln. of $Pb(OAc)_2$. The mixt. is dil. to a known vol., filtered and an aq. portion taken for the colorimetric detn. of the nitrite by means of *m*-phenylenediamine. The ppt. is washed with cold H_2O , suspended in cold boiled H_2O and the SO_3 detd. by acidification and distn. in a stream of CO_2 . Or, the sample may be dissolved in a dil. soln. of $NaHCO_3$ and the SO_3 detd. by the addn. of a slight excess of 0.1 N I soln. with back-titration with $Na_2S_2O_3$. The air is then removed by a stream of CO_2 and the soln. acidified with 1 cc. of 0.1 N H_2SO_4 , and after 1 min. without disturbance, the free I is titrated with $Na_2S_2O_3$. The stream of CO_2 serves to mix the 2 solns. KI should be added in the second titration only if the amt. of nitrite is in great excess. A max. error of approx. ± 1 mg. for both detns. was found in solns. of mixts. contg. 65-100 mg. of Na_2SO_3 and 25-100 mg. of $NaNO_2$. Meat samples are extd. with $NaHCO_3$, and the ext. should be free from substances which may interfere with the iodometric titration. C. R. FELLERS

The determination of the sulfite ion in the presence of the thiosulfate ion with the aid of fuchsin. V. E. MALINOVSKII *Ukrainski Khim. Zhurnal* 5, Sci. Pt., 181-6(1930).—Dissolve 0.34 g. fuchsin in 1 l. of water contg. 1 g. of H_2SO_4 , and heat slightly on a water bath. Make a satd. aq. soln. of the substance and dil. 16 times, add 1 drop of 1% phenolphthalein indicator and if the soln. is alk., carefully neutralize with 0.1 N $AcOH$. To 10-15 cc. of the neutralized soln. add a few drops of the fuchsin reagent. Decolorization within the first few min. is a certain indication of the presence of the SO_3^{--} ion. Under the conditions outlined the presence in the soln. of the SO_3^{--} ion does not interfere with the reaction. B. S. LEVINE

Determination of ammonium salts by the formaldehyde method. ERNESTO GILDO SORRENTINO *Atti III congresso naz. chim. pura applicata* 552-6(1930).—Five cc. of neutral 40% CH_2O and a few drops of phenolphthalein are added to 20-5 cc. of the NH_4 salt soln., and the mixt. is titrated with 0.2 N $Ba(OH)_2$ until a color persists. Results agree well with the theory. Urea, dicyandiamide and guanidine do not interfere. E. M. SMITH

The simultaneous test for thiocyanate and ferrocyanide anions by the drop method. A. V. PAVLINOVA AND T. N. BAKH *Ukrainski Khim. Zhurnal* 5, Sci. Pt., 233(German abstract 234)(1930).—Place a drop of $FeCl_3$ on a piece of filter paper, follow by a drop

of H_2SO_4 and finally by a drop of the soln. to be tested. In the presence of both anions a central blue spot of Prussian blue is formed surrounded by a red ring formed by the $\text{Fe}(\text{CNS})_3$. B. S. LEVINE

The effect of the nature of the paper upon the sensitiveness of the drop method. A. V. PAVLINOVA and T. N. BAKH. *Ukrainskii Khim. Zhurnal* 5, Sci. Pt. 235 (German abstract 236) (1930).—Ash free paper enables one to detect the presence of $\text{K}_2\text{Fe}(\text{CN})_6$ in $N/8445$ soln. while on ordinary paper the test is sensitive only to $N/5000$ soln. B. S. LEVINE

Poisoning by mercury oxycyanide. A. SARTORI. *Chem.-Ztg.* 54, 813-4 (1930).—An analysis of the viscera, performed 3 months after death, failed to detect the presence of HCN , but showed Hg . A. L. HENNE

Absolute determination of nitrogen in organic compounds. Pregl's micro-method. OTTO R. TRAUTZ and JOSEPH B. NIDERR. *Ind. Eng. Chem., Anal. Ed.* 3, 151-2 (1931).—The accuracy of the micro Dumas method for detg. N can be increased by using measured vols. of CO_2 and of CuO , applying empirical corrections for air and absorption errors allowing for adhesion to the walls and taking into consideration the vapor pressure of the KOH soln. over which the N_2 is measured. Detailed directions are given for carrying out an analysis with a few mg. of sample. Thus in one analysis of toluamide the value 10.33% N was obtained, which is correct within 4 parts in 1000 although the observed vol. of N_2 obtained was only 0.321 cc. and the calcd. corrections were 0.012 cc. W. T. H.

A micro-method for the determination of carboxyl groups in organic acids. SHUNJI TSURUMI and YASABURO SASAKI. *Science Repts. Tôhoku Imp. Univ.*, 1st Ser. 19, 481-8 (1930).—The carboxyl group in org. acids is detd. within an error of 1% on 2-5-mg. samples by a procedure based on that described by Hunter and Edwards (*C. A.* 7, 2189). The app. consists of a train, a decompn. vessel, an air chamber and a micro-buret (Flaschenträger, *C. A.* 19, 3230) in series. The train contains the usual elements for producing pure H_2S . The decompn. vessel (55 mm. \times 14 mm. diam.) has an inlet tube reaching the bottom, an outlet tube at the top and a side arm. A curved spatula, which may be rotated, is introduced through a stopper in the side arm. The air chamber is water jacketed and has 2 2.5-cc. bulbs and a 5-cc. bulb in series. Capillary tubing and stopcocks are used throughout. Charge the decompn. vessel with 1 cc. of carbonate-free NaOH and weigh the sample on a 7-mm. watch glass and support it on the spatula. Sat. the NaOH with H_2S from the train and shake to remove excess. Bring the system to atm. pressure and connect it to the buret. Spill the sample by turning the spatula, shake the decompn. vessel till no more gas is evolved and measure the displaced air over water in the buret. A. P. SHEPARD

Analysis of mixtures of hydrogen, methane and ethane. OSWALD J. WALKER and SOORYA NARAYAN SHUKLA. *J. Chem. Soc.* 1931, 363-70.—A Töpler pump is connected by means of a 3-way cock to a U-tube which can be cooled by immersion in a bath of liquid air for the removal of the C_2H_6 . It is also connected to bulbs containing "oxidized" Pd sponge for the removal of H_2 . The procedure consists, therefore, in condensing the C_2H_6 and measuring the contraction, adsorbing the H_2 by Pd and detg. the CH_4 by difference or by explosion. W. T. H.

Detection of glycerol. M. TRONCOSO SURANTES. *Rev. centro estud. agr. farm. biogum.* 18, 221-3 (1929).—Demigès' method (*C. A.* 22, 136) and Sanchez' procedure are modified by the use of BaO_3 instead of PbO_3 or MnO_3 , as oxidizing agent. B. C. A.

Reactions of glycerol and calcium glycerophosphate. LAD. ECKERT. *Pharm. Zentralhalle* 72, 85-9 (1931).—Various color tests are discussed in connection with the detection of glycerol, notably that involving the use of codeine and concd. H_2SO_4 . It appears unnecessary first to oxidize the glycerol to glyceraldehyde or dihydroxyacetone, since by simply putting a layer of concd. H_2SO_4 beneath an aq. soln. of glycerol and codeine a blue or violet color develops. An aq. soln. of Ca glycerophosphate shows like behavior. W. O. E.

Detection of isopropyl alcohol in brandy, spirits, tinctures, cosmetics and liniments by means of piperonal. G. REIV. *Z. Untersuch. Lebensmittel* 60, 213-54 (1930); cf. *C. A.* 23, 5542.—Twenty-five brandies, 40 cosmetics, 124 liniments and various samples of spirits, tinctures, etc., were examd. by the following improved technique: Ten cc. of the sample is distd. on the water bath into a small cylinder immersed in ice water until no more alc. is evolved, 0.3 cc. of the distillate is shaken with 0.7 cc. of a mixt. of 80 cc. of abs. alc. and 20 cc. of H_2O , and a soln. of 0.1 g. of NH_4OH in 3 cc. of water is added. The mixt. is shaken for 3 min. at room temp., 0.4 g. of absorbent charcoal is added and the mixt. is shaken and filtered through a small dry paper into a 100-cc. round-bottomed flask. Five cc. of a freshly prepd. soln. of 0.5 g. of piperonal in 100 cc. of abs. alc. is added,

followed by the slow addn. of 20 cc. of H_2SO_4 , sp. gr. 1.84. The mixt. is shaken and placed on the water bath for 5 min. in a large test tube, a red color indicates the presence of isopropyl alc., and a green brown color its absence. If 30 cc. of 30% AcOH is added, a stable rose-red color is a pos. reaction, a brown color indicates the presence of a small quantity of the alc., a colorless soln. or weak red tinge stable for a min. and turning yellow gray or colorless shows a definite reaction. The sensitiveness of the reaction is 1-2% for most substances but is delicate to 0.1% with liniments. In the latter, CH_3O if present, must be removed before distn. by heating the sample with 5 cc. of NaOH soln., sp. gr. 1.12, under a reflux condenser for 1 hr. Concns. of acetone less than 20% have no harmful effect on the reaction. C. R. FULLER

Colorimetric determination of the tyrosine and tryptophan contents of various crude protein concentrates. WM. D. McFARLANE AND HUGH L. FULLER. *Biochem. J.* 24, 1601-10(1930).—The detns. were based on the method developed by Folin and Ciocalteu (*C. A.* 21, 3210). The tyrosine (ty) and tryptophan (tr) contents in one series of expts. were as follows:

	Ty %	Tr %
Casemogen	7.19	1.54
Buttermilk powder	6.36	1.25
Meat meal	3.02	0.95
Tankage	2.42	0.83
Cod liver meal		0.77
Fish meal	2.92	1.38

BENJAMIN HARKOW

The separation of Be and Al oxides (TRAYERS, SCHNOUTKA) 6. Technic of extracting liquids that form emulsions (WEINBERGER) 2. Apparatus for electrical tests to identify various materials (U. S. pat. 1,793,970) 4.

JOLINAUX ALCEIDE. *Leçons de chimie analytique*. Paris. Hermann et Cie. 350 pp. F 60.

MERCK E. *Prüfung der chemischen Reagentien auf Reinheit*. 4th ed. Darmstadt. Merck. 401 pp.

Determining monatomic substances in gases, particularly mercury vapor in air. SIEMENS & HALSKA A. G. (Otto Krenzien, inventor). Ger. 517,480, Sept. 13, 1924. The method depends on the detn. of the intensity of the resonance effect produced when a beam of light, the spectrum of which includes a "resonance line" of the spectrum of the element to be estimated, is directed through the gas in which the element is to be detd. and then into a vessel also contg. the vapor of the element.

8—MINERALOGICAL AND GEOLOGICAL CHEMISTRY

EDGAR T. WHERRY AND J. F. SCHAIKER

The formation of calcite crystals by slow precipitation of calcium salts in water by ammonia. J. HENKOW. *Ann. soc. géol. Belg.* 53, B40-1(1929).—A flask of tap-water connected by an overhead glass tube to a flask of concd. NH_3 pptd. calcite, identified petrographically on 4 months' standing. R. H. EWELE

The precipitation of calcium carbonate in dilute solutions. F. CORRY. *Ann. soc. géol. Belg.* 53, B64-6(1930).—Referring to Henkion's paper (preceding abstr.), C. states the present status of the subject, and cites the work of several earlier investigators. R. H. EWELE

The deposit of cerite in Kystym district. V. ZILBERMANTZ. *Trans. Inst. Econ. Mineral.* (Moscow) No. 44, 41-2, & *C. A.* 24, 4243.—Assoc. with parashite there has been found another mineral, differing in all its properties but closely connected genetically. It contains about 19% SiO_2 , 11% CaO and 64% Ce, La and "Di" oxides. It approaches cerite in optical properties and compn. but is purer than Swedish cerite. H. C. PARKER

Radioactive mineral found in Japan. SATOYASU IDOMI, JUN YOSHIMURA AND SHIN HATA. *Sci. Papers Inst. Phys. Chem. Research* (Tokyo) 15, No. 285, 83-8(1931).—A radioactive mineral has been found in the Bijozan Range, in black prismatic crystals about 6-7 mm. long, with a resinous luster and a hardness of 5.5. The powder and the streak are brown. It is easily decomposed by HCl, leaving a white flocculent residue.

It is a new species belonging to the epidote group, and the name *nagalelite* is suggested. Its radioactivity is thought to be due to a small quantity of Th. An intensively radioactive mineral was found in a biotite crystal. It gave a qual test for U. Laterite and lateric clays were also found. Analyses of some of the minerals are given.

L. D. ROBERTS

A variety of uranotile from Chinkolobwe (Katanga). J. THOREAU *Ann. soc. géol. Belg.* 53, B60-4(1930).—A sample of pitchblende was encrusted with a thin layer of yellow radiating needles with black opaque inclusions, weakly dichroic, with parallel extinction, + elongation, $n_x = 1.703$, $n_z = 1.632$. The d is greater than 3.3. Analysis gave volatile matter (H_2O) 15.4, insol. matter (SiO_2) 6.8, UO_2 73.1, CaO 7.4%. This corresponds roughly to $SiO_2 \cdot 2UO_2 \cdot CaO \cdot 6H_2O$. The mineral dissolves in cold concd. HCl with evolution of a gas, and if heated the black inclusions also dissolve. The black mineral may be uraninite. There is also another yellow, fibrous mineral on the pitchblende with somewhat different properties (n_z less than 1.676), which may be another variety of uranotile.

R. H. EWELL

Sorption of gas by minerals. II. Laumontite. JITSUSABURO SAMESHIMA *Bull. Chem. Soc. Japan* 5, 303-10(1930).—Laumontite ($CaAl_2Si_4O_{12} \cdot 4H_2O$) sorbs NH_3 , but not CO_2 . The sorption data are tabulated and presented in curves. The dryness of the mineral has an important influence. NH_3 combines chemically with the dehydrated mineral.

ALFRED L. HENNE

The presence of apatite in the Upper Cambrian rocks of Neuville. Note on the optical study of weakly birefringent minerals. F. CORRY, *Ann. soc. géol. Belg.* 53, B66-8(1930).—After considerable difficulty, C. identified apatite as an accessory constituent of these rocks. The birefringence is only 0.0015.

R. H. EWELL

Contribution to the mineralogy of the coal fields. M. BELLIERE *Ann. soc. géol. Belg.* 53, B125-39(1930).—The mode of occurrence in various coal mining regions of Belgium is described for the following minerals: S, sal ammonite, ankerite, quartz, pholentite, pyrite, chalcocopyrite, sphalerite, galena, calcite, dolomite, gypsum and aragonite. Crystallographic data and chem. analyses are given.

R. H. EWELL

Petrographical description of the Sadon mine. E. A. KUZNETZOV *Trans. Inst. Econ. Mineral. (Moscow)* No. 46, 3-53(1930).—The apatite veins are adapted to a transverse system of cracks, which are parallel to a more recent thrust. The same direction was followed by keratophyres at a later period. After a long period of rest accompanied by the destruction of granite and deposition upon it of sedimentary rocks, new movements produced new cleavages, along which flowed the overlying keratophyres. Then, after another period of relative rest, a new fracture occurred in the same region, where granite, porphyry and sedimentary formations were caught by the movement and ground up. Simultaneously ore-bearing solns. rose from their original depth reservoir. The granite was crystal. under the influence of its own gas phase. Apparently the hydrothermal solns. were rather hot at first. Later they cooled and the main manifestation of their force is to be seen in silicification, sericitization, chloritization and carbonatization of adjacent rocks.

H. C. PARISH

Mineralogy of the Sadon ore vein. E. E. ZAKHAROV, *Trans. Inst. Econ. Mineral. (Moscow)* No. 46, 54-152(in English) 153-4(1930).—The Sadon deposit of Pb-Zn ore showed galena, sphalerite, pyrite, chalcocopyrite, pyrrhotite, arsenopyrite and molybdenite, of the gang minerals quartz, carbonates, chlorite, muscovite and a mineral of the gummites series. A series of the minerals smithsonite, cerussite, malachite, azurite, gypsum and Pb oxide has also been discovered. Z. notes breccia structures, imbricated and banded structures, a thick network of ore-bearing veinlets, a thick ore inclusion and structureless formations.

H. C. PARISH

Magnetite: its application in assaying. R. J. B. KETHEL, *J. Chem. Met. Mining Soc. S. Africa* 31, 172-5(1931), cf. C. A. 24, 5256.—Discussion. ALDEN H. EMERY

Asbestos deposits in the Union. A. L. HALL, *S. African Mining Eng. J.* 41, II, 423-4(1930), 487-8(1931).—The Union of South Africa contains asbestos deposits of very considerable extent. It is the leading crocidolite producer. The following fiber varieties occur in the Union: amosite, crocidolite, chrysotile, anthophyllite, tremolite. The occurrences of each are described.

ALDEN H. EMERY

Petrology of chrysotile asbestos deposits of the Krasno-Uralsky asbestos mine in the Ural Mountains. V. V. ARSHINOV AND B. YA. MERENKOV, *Trans. Inst. Econ. Mineral. (Moscow)* No. 45, 1-83(1930).—The asbestos is graded according to fiber length as follows: grade I about 16 mm, grade II, 12 mm, III, 8.5 mm; IV, 4 to 2 mm; VI, 2 to 0.5 mm. It is believed that the serpentines of this region have been derived from saxonites under the simple influence of water. Other basic rocks encountered in small quantities are schlieren of pyroxenite, a mono-mineral rock formed by diopside.

The serpentinization of saxonites is developed along the fissures as ribbons of serpentine. As this process goes on, the grain size diminishes until the whole rock is altered into serpentine. The characteristic types of asbestos ribboning are represented by the following varieties: (1) bordered asbestos; (2) fine asbestos veinlets branching; (3) fine asbestos parallel sets. A. believes all of these rocks to be products of the differentiation of diorite intrusive bodies. The absence of granite pegmatite dikes is characteristic for the asbestos deposits of the Krasno-Uralsky mine. H. C. PARISH

A new deposit of plastic clay. L. NYS *Ann soc. géol. Belg.* 53, 877-8 (1930). The deposit is at Grèves Perres. Analysis gives: SiO_2 73.08, Al_2O_3 19.97, Fe_2O_3 1.78, CaO 0.10, MgO 0.06, loss on heating 4.30%. There are inclusions of yellow sand and small amounts of andalusite, zircon, magnetite and tourmaline. R. H. EWELL

Studies on the moisture adsorbed by the Kanbara clay. HAJIME ISONE *Sci. Papers Inst. Phys. Chem. Research (Tokyo)* 14, 223-74 (1930) (in English).—In certain deposits of Kanbara clay the upper layers are acid, the middle layer is neutral and the lower alk. Neutral clay can be prep. from either the acid or the alk. clay by washing with dist. water. Acid and alk. clays show different vapor pressure curves, but the curves of the natural neutral clay, clay artificially neutralized with acids or alkalis and clay made neutral by washing with dist. water, have practically the same characteristics at the same temp. The water adsorbed by special capillary action ranges from 20 to 35%, while that due to true adsorption of the surface of the clay is about 2 to 18%.

On the basis of x-ray studies the formula $[\text{Al}(\text{SiO}_3\text{SiO}_3)_2][\text{Mg}]_2\text{H}_2\text{O}$ is advanced for the clay. H. K. D. JACOB

The transparent and opaque diamonds of the Bushma deposits. EDMOND POLINARD *Ann soc. géol. Belg., Publ. rel. au Congo Belg.* 52, C179-218 (1929); The transparent diamonds of the Bushma deposits. *Ibid.* 53, C1-45 (1930).—Complete and detailed crystallographic data are given. R. H. EWELL

The igneous rocks of the Helle (Haute Fagnes). PAUL RONCHESSE *Ann soc. géol. Belg.* 54, B35-47 (1930).—Petrographic analysis shows these rocks, formerly classed as granites, to be quartziferous diorites because they contain about 30% of anorthite in the plagioclase and a small amount of orthoclase. Analysis gives: SiO_2 70.06, Al_2O_3 14.86, Fe_2O_3 1.64, FeO 0.63, FeS 1.28, MnO 0.09, CaO 3.45, MgO 1.02, K_2O 2.06, Na_2O 4.83%. Accessory minerals include biotite, chlorite, magnetite, epidote, apatite, zircon, scapolite, laumontite and sulfides. R. H. EWELL

The plagioclases of the igneous rocks of the Helle. P. MICROT *Ann soc. géol. Belg.* 54, B41-8 (1930). cf. preceding abstr.—M. discusses the mutual soly of orthoclase, albite and anorthite to explain the absence of separate crystals of orthoclase, when its presence is shown by chem. analysis. The zoning of the plagioclases is explained by the effect of irregularly varying pressures on the temp.-compn. diagram of the system. R. H. EWELL

The pegmatites of the sodalitic syenite of the Island Roume (archipelago of Los, French Guinea). Description of a new mineral (serandite) found there. A. LACROIX *Compt. rend.* 192, 169-94 (1931).—The island is very rich in feldspathoid syenite. Mineral constituents of the pegmatites are described, and mineralogical characteristics of normal syenite are compared with those of the pegmatites. The new mineral called serandite is described in detail and chemical analysis given: SiO_2 48.72, Al_2O_3 0.29, Fe_2O_3 0.03, FeO 1.83, MnO 28.99, MgO 0.06, CaO 10.42, Na_2O 7.38, K_2O 0.20, $\text{H}_2\text{O} + 2.67$, $\text{H}_2\text{O} = 0.11\%$, corresponding to SiO_2 (MnO , CaO) (NaO K_2O) H_2O . A somewhat similar manganese bearing pectolite from Magnet Cove, Ark., is discussed, but the Mn is in smaller quantities, and d and optical properties are different from those of the new mineral. ALICE W. FERRISON

The alkaline rocks of the southern part of Kystym district, Ural. E. A. KUZNETZOV *Trans. Inst. Econ. Mineral. (Moscow)* No. 48, 1-89 (1930).—This area is formed of granite gneisses, syenites and muscovites. In the granite gneisses are bands of amphibolites, quartzites and garnet-muscovite gneisses. Analyses of the muscovites are given. On the mountain Sobachy all the muscovite bands are edged by alk. syenites forming the reaction zones of muscovite and granite gneisses. The western dikes of muscovites are more or less sharply sep. from the country rocks, the latter being altered into syenite gneisses with aegirine, augite and biotite. As against the Eastern band the muscovites become fine grained in contact and at some places contain considerable porphyritic inclusions and albite or claurolite up to 5 cm. In other respects all muscovites are similar. H. C. PARISH

Manupolite and its related rocks. J. MOROZEWICZ. *Prace Polsk. Inst. Geol.* 2,

217-350(1929). *Tschermak's Mineral. Petrog. Mitt.* 40, 335-436(1930) —Mariupolite, from Ukraine, is an extreme member of the nepheline-syenite series. It consists of albite, nepheline and aegirine, with sometimes lepidomelane, sodalite or cancrinite, and accessory magnetite, beekelite and pyrochlore. The order of intrusion of this and the associated rocks was that of increasing basicity.

	Mean d	Mean SiO ₂ (mol.-%)
1 Upper biotite granite	2 632	80 0
2 Diablage and amphibole-granite	2 678	74 4
3 Alkali syenites	2 699	70 5
4 Mariupolites	2 712	63 9
5 Foyaites	2 718	60 6
6 Wehrlite	3 359	46 2

Many chem. analyses are given of these rocks and their constituent minerals, and the relations are discussed in detail.

Origin of certain Pre-Cambrian amphibolites in Agder (Southern Norway). TOM F. W. BARTH. *Norsk Geol. Tids.* 11, 219-31(1930) —Pre-Cambrian amphibolites found in Agder have the same 3 modes of origin as those found by Adams (*Compt. rend. Congr. Géol. Int. Stockholm* 1910, 536-72) for amphibolites from Haliburton and Bancroft areas (*Memoir 6, Can. Geol. Survey*, 1910). Comparative compns. are shown. A para-amphibolite formed by metamorphic action of granitic magma or "ichor" on impure calcareous sediment embedded in the magma is described. A map of a small area of Pre-Cambrian rocks in which amphibolites originated by alteration of limestones by the action of granitic "ichor" is shown. The same mechanism produced an amphibolite which formed a basic schlieren in the anorthosite on the eastern slope of Mt. Pokamoonshine in the Adirondaeks. Detailed data on different amphibolites are presented. Actinolitic hornblende occurs in Norwegian amphibolites formed by alteration of pure limestones, other types contain common hornblende. A theory of formation is discussed.

Structure and form of fresh-water limestones. E. ROZENŠTEINS. *Acta Univ. Latviensis Kim. Fakultat. Seriya 1, Fas 1* Burtneka, (in German 13-21) (in Lettish 25-7(1929)) —The compact limestones consist of microcryst (0.01-0.2 mm) or mesocryst (0.25-0.07 mm) calcite aggregates in rosette or druse forms. The medium compact limestones are of crypto- and pelitomorphous aggregates (0.003-0.005 mm) and often the interspaces are filled with secondary crystals. Around thin plant fibers there forms a strong crust which tends to form in layers. The gypsum content may range from 1.2% to 4.12% and may be found in limestones deposited from gypsum-free water, on account of the oxidation of pyrite. Silica was as high as 13.02%.

Structure of weather-resisting rocks. A. L. W. E. VAN DER VEEN. *1st Communications New Intern. Assoc. for Testing of Materials B*, 10-2(1930) —Fossilization (cementation) of a sediment is made by solns. circulating in its interstices. The microscope reveals the mineralogical compn. of the sediment and the nature of its cement. The chem. analysis of a weather-resisting rock may suggest that something is wrong within, whereas the microscope reveals that the dangerous atoms are locked up in most resisting minerals. The relation between genesis and weather resistance of some sandstones and marbles is described.

Mineralogical study of the soil of the Vercelli district. E. REPOSSI. *Atti accad. sci. Torino Classe sci. fis., mat. e naturali* 64, 335-58(1929).

B. C. A.

The ceramic industry of Ontario (MONTGOMERY) JR. *Am.* —its geographic distribution and geological occurrence (BELL) 9. Phase equilibrium in binary systems with continuous mixed crystals (KORDEZ) 2. Gravity separation (EMMONS) 2. The geochemistry of the Ti group (VON HEVESY) 2.

9—METALLURGY AND METALLOGRAPHY

D. J. DEMOREST, H. W. GILLET and RICHARD RIMBACH

Present status of the dressing of bituminous copper ores. A. GÖTTE. *Metall. u. Erz* 28, 73-5(1931) —The most notable deposits of bituminous Cu ores in Germany are at Niedermansberg and Mansfeld. A discussion of dressing methods used for these ores is given. The bituminous content in no way interferes with the flotation processes.

H. STOKERTZ

Investigation of ore dressing machines with the help of yield calculations. W. LUCKEN AND L. KRAEPPE *Metall u. Erz* 28, 49-53 (1931).—A study of operating conditions at the "Ant" mine in Siegerland by comparison of the yields obtained from a dust settling machine and a Hercules hearth. The particle range is 22-2.5 mm for the first machine and 2.5-0 for the second. H. STOERTZ.

Effect of particle size on flotation. A. M. GAUDIN, JOHN O. GROW AND H. B. HENDERSON *Am. Inst. Mining Met. Eng. Tech. Pub. No. 414*, 3-23 (1931).—A study is made to determine what effect particle size has on floatability. One method of study was on the samples of the final products of several mills and in the other mixtures of pure minerals were ground to different degrees of fineness and floated. In both cases the products were sized and analyzed. Conclusions: The results obtained by the 2 methods are in agreement. The usual notion about the relative ease of flotation of very fine particles does not agree with the facts. There is a time sequence of flotation with regard to size, the medium coarse particles floating first. No successful method of floating extremely fine pulps has been found. Recovery is optimum in a well-defined size range. Selection is optimum in another well-defined range. The bulk of the pulp is in the range of 100-800 mesh. Thirteen figures and 5 tables are shown. Data from lab and practice agree closely. W. H. BOYNTON.

Experimental flotation of oxidized silver ores. H. S. GIESSE *Am. Inst. Mining Met. Eng. Tech. Pub. No. 401*, 9 pp (1931).—Because of the other uses the use of organic compounds available for flotation is limited. Benzyl mercaptan and benzyl sulfide and disulfide show interesting results. Thioureas leave high tailings. When the diethylphenyl and di-*o*-tolyl thioureas are used the froth is white and barren. Mercapto-benzothiazole is of some value with a number of oxidized precious metal ores. Thioacetanilide gives fair results, while thio-benzanilide is better. Reagent soly. is an important factor in flotation. H. C. PARISH.

Flotation of minor gold in large-scale copper concentrators. E. S. LEAVER AND J. A. WOOLF *Am. Inst. Mining Met. Eng. Tech. Pub. No. 419*, 32 pp (1931).— CaO is an active depressant for free Au during flotation if the alkyl of the soln circuit is $p_n = 10$ or above. A reaction of $p_n 7.5$ to 10 is best. In the flotation of base metal ores containing Au as a by-product careful regulation of CaO , avoiding excess, will increase Au recovery. If pyrite in the ore is crushed fine enough to free the Au, most of the pyrite may be depressed by low CaO alkyl in the flotation circuit and a high percentage of the Au floated with the Cu sulfides. Ores with a low ratio of pyrite to CuS and containing under \$1 in free Au per ton gave, by control of the flotation circuit assuring low alkyl, 25% Cu concentrates containing more than 90% of the Cu and more than 80% of the Au. Ores high in pyrite and lean in Cu sulfides and Au as a nominal by-product, after low grade concentrates unless the percentage of Au recovery is sacrificed. In the flotation of oxidized base-metal ores in which a sulfidizing reagent is used, Na_2S gives a much higher Au recovery than CaS or Ca polysulfide , Ca salts depress Au. H. C. PARISH.

Investigation on the influence of the iron content of zinc blendes on their floatability. LUDWIG KRAEPPE *Mitt. Kaiser-Wilhelm Inst. Eisenforsch. Düsseldorf* 12, 343-52 (1930).—Flotation tests were made on blendes in a mechanically agitated cell and in a pneumatic cell. It was found that each sample of blende gave a definite p_n to the water regardless of whether the water was originally slightly alk. or acid, the value being lower for blendes of high Fe content. The max. recovery was obtained with a p_n between 6 and 7, and it appeared that the optimum p_n was lower for high Fe content. Very small quantities of CuS speeded up the rate of recovery. The quantities of KCN and ZnSO_4 necessary to hold down the Fe differed widely for different ores but bore no simple relation to the Fe content. Conclusion: The presence of foreign materials as well as the pretreatment is of more significance than the Fe content in affecting floatability. H. C. DREUS.

Quinquennial review of the mineral production of India for the years 1924 to 1928. *Records Geol. Survey India* 64 (1930).—Gold. G. D. P. CORRIE *Ibid.* 92-106.—Indian Au occurs both in veins and as alluvial deposits. Individual deposits are described. Production has fallen steadily since 1915. Lead. G. VERNON HOBSON *Ibid.* 154-66.—Near Bawdwin ore bearing solns have ascended through crushed rhyolitic tuffs, flows and breccias metasomatically replacing materials in the tuff and leaving sulfides. The ore is an intimate mixt. of Ag bearing galena and sphalerite and in many places also chalcopryite. Production for 1924-28 averaged about 60,000 tons of Pb and 5 1/2 million oz Ag. Manganese. L. L. FERNOR *Ibid.* 172-233.—Indian Mn occurs chiefly in metamorphosed manganeseiferous sediments containing spessartite and rhodonite. Typical ores contain braunite and palomelane. Ores are also found in association with cryst. limestone containing pyromorphite. Other minerals found are holland-

lte, vredenbrigitte, sitapelite and rarely pyrolusite. The high grade ores contain Mn 48-54, Fe 4-8 and SiO_2 6-9%. World production, analyses of ores, manu. of ferromanganese in India, various economic considerations, etc., are discussed. India is the world's leading producer. **Silver.** G. VERNOV HONOV *Ibid* 213-8.—No ores are mined primarily for Ag in India. Over 99% of the production is a by-product from Pb, the remainder from Au. The government Ag refinery is described in detail. **Tin.** J. COGORY BROWN *Ibid* 208-303.—Almost all India's Sn production comes from Burma cassiterite, the occurrence of which is described in the section on tungsten. **Tungsten.** J. COGORY BROWN *Ibid* 303-10.—Wolfram and cassiterite occur in quartz veins closely associated with a biotite boss-granite intruded into a series of slates, shales, argillites and agglomerates with greatly subordinate quartzites, limestones and conglomerates. The veins fill fissures. The mineral associates of wolfram are not the same in different deposits. The various associated minerals are beryl, tourmaline, mica, fluorite, molybdenite, pyrrhotite, galena, sphalerite, arsenopyrite, Bi, bismuthinite and topaz. The vein formation appears to be a direct sequence of processes of differentiation influenced by local conditions. **Zinc.** G. VERNOV HONOV *Ibid* 310-2.—Sphalerite occurs with galena at the Dawdwin mine. The industry is growing.

ALDEN H. EMERY

The composition of Finnish ores. GUST A. AARTOVAARA *Yhteinen Tekn. Julkaisulehti* (Reprint) 10, 1-9(1927).—Analyses are given of sulfide and oxide ores.

J. RYSELYN

The use of Greek nickel ores in Germany. E. REITLER *MetallZest* 20, 2270-4 (1930).—Before the war the German nickel works used New Caledonian ore and some American oxide. Attempts have been made to use Greek nickel ores from the Laurion Mountains. The composition of these ores is, on the dry ore: NiO 6.62, CoO 0.15, Fe_2O_3 40.01, SiO_2 28.83, Al_2O_3 14.61, As_2O_3 0.04% and minor quantities of CuO , CaO , MgO and Cr_2O_3 . The smelting of these ores is complicated by the presence of so much alumina, which give very viscous slags. It is very doubtful if these ores can be successfully used.

R. S. DEAN

Iron ore on Canyon Creek, Fort Apache Indian Reservation, Arizona. ERNEST F. BURCHARD. U. S. Geol. Survey, *Bull* 621-C, 51-75(1931).

E. H.

Rationalization of the fusion of lead minerals. N. C. KYRIACOU. *Génie Civil* 93, 194-9(1931).—Insufficient prep. of the mixing bed and imperfect combustion cause the consumption of coke. The metal should be dense, free of S, in the form of the silicate, perfectly clinkered and free of dust. The most suitable size of granule is 20 mm in diam. A rotary furnace for preliminary roasting feeds its charge directly to the shaft furnace. The use of a water jacket and of a hearth reduces Pb losses.

ANN NICHOLSON HIRD

Treatment of low-grade gold ores at West Rand (South Africa). JEAN FULOR. *Génie Civil* 93, 153-8(1931).—Modern equipment has raised the Au recovered by the cyanide process from 49.78 to 61.93%. The Au in the tailings is lowered from 0.316 to 0.256 pennyweight per ton.

ANN NICHOLSON HIRD

Origin of flour gold in black sands. A. E. KELLOGG. *Mining J* (Arizona) 14, 3-4, 49-50(1931).—Flour Au may float over riffles, adhere to black sand or magnetic Fe or refuse to amalgamate. Flour Au may have been originally minutely dispersed through rocks, run into solution with quartz in vein fissures, wrapped up within or chemically combined with other Au ores or may have been formed by attrition by waves. An example of such grinding is in the very ancient Siskiyou Island. No reliable means of recovering flour Au is known.

B. M. SIMMES

Gold—its geographical distribution and geological occurrence. J. MACINTOSH BELL. *Canadian Mining J.* 52, 216-51(1931).—An address is given which covers the world's gold fields. It seems probable that the annual production of Au will decrease 20-25% during the next decade.

W. H. BOYNTON

Mining and mill in the Vermont copper district. C. S. ANDERSON. *Eng. Mining J.* 131, 208-10(1931).—Mining of the cupriferous pyrrhotite deposits of Vermont is outlined. The Cu content varies between 0.5 and 20% with an av. of about 3%. A flowsheet is shown for the Elizabeth-Foster-Cleveland mill of the National Copper Corp.

W. H. BOYNTON

Milling methods and costs at the Spring Hill concentrator of the Montana Mines Corporation, Helena, Montana. L. A. GAAYE. *Bur. Mines, Information Circ.* 6411, 8 pp (1931).—The ore, Au (10-15% native) associated with pyrite, arsenopyrite and Bi, occurs disseminated through diorite along a limestone contact. Barren pyrrhotite, the most plentiful sulfide, floats with the Au-bearing sulfides. Costs per ton of ore mined for 1929-30 were crushing and conveying \$0.102, grinding \$0.290, flotation

\$0.258, tailings disposal \$0.032, H_2O \$0.012, repairs \$0.233 and total \$1.007.

ALDEN H. EMERY

Milling methods and costs at the lead-zinc concentrator of the Treadwell Yukon Company, Limited, at Tybo, Nevada. W. H. BLACKBURN. *Bur. Mines, Information Circ.* 6430, 14 pp (1931).—The ores are complex and contain Pb, Zn, Fe, Ag and Au. Most of the Ag is associated with galena. Sphalerite and pyrite are also argentiferous. Pb concentrates average 90.43 oz Ag per ton, 62.75% Pb and 3.50% Zn. Zn concentrates average 17.05 oz Ag per ton, 2.71% Pb and 46.57% Zn. Costs for March, 1930, were in cents per ton of ore treated: sorting 1.8, crushing 7.9, grinding and classification 44.5, flotation 50.2, filtering 2.1, general 21.7 and total \$1.282. Pb recovery is 74%, Zn 70%, Ag 75%. Detailed metallurgical data are given. ALDEN H. EMERY

The absorption of gold in tube mills. J. A. WHITE. *J. Chem. Met. Mining Soc. S. Africa* 31, 161-71 (1940).—Though concn upon the surfaces of pebbles and tube linings is a real effect, it does not account for more than a minor portion of the total Au locked up in the tube. The remainder is held in the numerous joints in the linings, between the lining and the shell and in other crevices. ALDEN H. EMERY

Amalgamation practice at Porcupine United Gold Mines, Ltd., Timmins, Ontario. RONALD A. VARY. *Bur. Mines, Information Circ.* 6433, 5 pp (1931).—The ore consists of quartz with fine stringers of mineralized basaltic schist, or of banded quartz schist, or of stringers of quartz in schist, accompanied by pyrite and often contg. visible Au. It is amalgamated in a 2½-ton pilot mill. Costs per ton of ore treated are crushing \$0.341, grinding \$0.771, classifying, screening, conveying and refining \$0.626, miscellaneous \$0.070 and total \$1.831. ALDEN H. EMERY

Historical evolution of Lower Hartz smelting methods. ROSENHAIN, *Mettall. u. Erz* 28, 75-85 (1931). H. STOEHR

The rationalization of zinc smelting in horizontal retorts. H. BREMANN. *Mettall. u. Erz* 20, 2365-6, 2441-2 (1930).—The slowness of the reaction is due to the inability to carry it on at a very high temp. since the reaction zone proceeds inward concentrically from the retort walls and consequently the Zn reduction reaction is going on at a gradually decreasing temp. and hence rate. From these considerations B. concludes that a small quantity of some gas which would not interfere with condensation could be profitably introduced into the retort to equalize the temp. somewhat. No test has been made of the idea. R. S. DEAN

The reaction between magnetite and ferrous sulfide. H. F. S. WARTMAN AND G. L. OLDSHOTT. *Bur. Mines, Rept. of Investigations* 3072, 10 pp (1931), cf C. A. 23, 1374.— SiO_2 and MgO accelerate, whereas CaO , Al_2O_3 and Cu_2S hinder, the reduction of magnetite by FeS . Orthoclase, anorthite, albite and kaolin also accelerate the reaction. ALDEN H. EMERY

Nickel slagging in copper refining with the formation of ferrites. W. STANL. *Mettall. u. Erz* 28, 53-6 (1931).—In the presence of Fe it is comparatively easy to remove Ni from Cu by slagging it out as ferrites ($NiFe_2O_4$ and $NiFeO_4$). The residual Cu still has a Ni content of 0.10-0.15% which cannot be removed in this manner. H. S.

Suggested improvements in fire refining of copper. H. H. ALEXANDER. *Eng. Mining J.* 131, 226-9 (1931).—An outline is given of methods incorporated in U. S. Patents 1,687,277 (C. A. 22, 4450), 1,756,967 (C. A. 24, 3207) and 1,776,826 (C. A. 24, 5721). Means of removing impurities by oxidation, the deoxidation of Cu and improved melting practice as tried out in the lab. and in furnaces of com. size are pointed out. Melting and refining simultaneously result in low initial cost of equipment, lower slag production, few repairs, low operating cost and a better product. W. H. ROYNTON

The zinc industry in Belgium. V. FICKET. *Bull. soc. belg. eng. ind.* 10, 799-844 (1930). E. H.

Photoelectric control for soaking-pit covers. R. M. BAYLE. *Elec. World* 97, 409 (1931). C. G. F.

Fluorspar in the open-hearth slag. H. L. GRIGER. *Blast Furnace Steel Plant* 19, 412-4 (1931).—A survey of the work of prominent metallurgists to help to make clear the much discussed function of fluorspar in the thinning of slag. Basic open hearth tapping slags do not contain F. F is probably lost as volatile SiF_4 , and a little as a volatile F-S compd. (In the basic elec. process very little F is lost, however, very little SiO_2 is present.) The action of fluorspar is probably to decrease the SiO_2 in the slag by the formation and loss of SiF_4 . In order to establish equi. FeO and Fe_2O_3 from the bath replace the lost SiO_2 . It is generally agreed that these metallic oxides increase fluidity. If this analysis is correct, then Fe oxide is the actual thinning agent. A. W. HOLMES

The economic situation of non-ferrous metallurgy. LÉON GUILLET. *Génie Civil* 98, 81-5(1931).—International statistics are given on the production, export and import of Cu, Pb, Zn, Sn, Sb, Ni, Al, Hg, Ag, Pt and Au from 1870 to 1928. A. N. II

X-rays in the foundry trade. R. BERTHOLD. *Gießerei-Ztg* 26, 642-4(1929).—The economic value of the use of x rays in the foundry is discussed with particular reference to light metal castings. MALCOLM DOLE

X-ray technic in the foundry. WALTER D. BARTELS. *Gießerei-Ztg* 26, 632-4(1929).—A discussion of the technic, costs and profits of the use of x rays in the foundry. MALCOLM DOLE

Rosin in core oil. WERNER G. SMITH. *Proc Am Soc Testing Materials* 30, Pt. II, 839-41(1930).—The compn of core oils and their use in the foundry are discussed. During baking of the core much of the core oil is lost by volatilization, leaving a porous but strong material which is free from gas producing substances. Rosin imparts tensile strength to the oil film and prevents excessive absorption of the oil by sands contg loam. The medium grades of rosin are commonly used for this purpose. S suggests improvements in the packaging and shipping of rosin. H. K. SALZBERG

The influence of colloidal ferric hydroxide upon the properties of molding sand. ETUJI TAKAHASHI. *Kinzoku No Kenkyu (J for Study of Metals)* 7, 552-62(1930).—A difficulty connected with synthetic molding sands consists in the lack of durability that is, fine clayey substance is liable to sep from the sand. In a study of the action of electrolytes and $Fe(OH)_3$ sol upon clay substance in molding sand, T found that the addn of colloidal $Fe(OH)_3$ increases either the plasticity or the gas permeability according to ramming conditions. The increase in gas permeability seems to be due to the coagulation of clayey particles. It prevents the separation of bonding substance from the sand. Various applications of colloidal $Fe(OH)_3$ for making synthetic sands were also studied. M. KURODA

Influence of siliceous matters upon the reduction of magnetic sands. KEIZO IWASE and MASAJI FUKUSHIMA. *Kinzoku No Kenkyu (J for Study of Metals)* 7, 524-34(1930).—The influence of siliceous matter such as tailings, sea sand, pptd silica and alumina upon the rate of reduction of magnetic sands has been detd at temps between 900° and 1150°. At lower temps their influence is small because no sintering occurs but at higher temps the semi reduced product FeO from the hematite in the tailings and other substances combine with silica forming semi fused substances and retard considerably the rate of reduction. M. KURODA

The manufacture and properties of malleable castings. J. K. SMITHSON. *Foundry Trade J.* 44, 110-12(1931).—A description. H. C. PARISH

New methods of producing quality gray cast iron, hard cast iron and tempered cast iron in copola furnaces. LAMLA. *Gießerei-Ztg* 27, 621-3(1930).—A general discussion. H. C. DUVS

Low-carbon, malleable cast iron from cupola furnaces. W. VALENTIN. *Gießerei-Ztg* 27, 617-22(1930).—A general discussion. H. C. DUVS

Advantages of fine-grained iron over coarse-grained iron for the production of high-grade castings. R. SCHMITT. *Z ges Gießereipraxis* 51, 25-8(1930); *Chimie & Industrie* 25, 8t(1931).—By the use of fine-grained iron instead of hematite iron, castings with identical C, Si, Mn, P and S contents but of finer grain were obtained, the tensile strength increased from 23 to 26 kg per sq mm, the Brinell hardness was unchanged, and the mech properties of the castings were improved in spite of an increase in the graphitic C content. The effects of the size of the grains were evaluated quantitatively, and variations in the mech properties resulting from the metallographic tests confirmed the following mechanism. When fine-grained iron is used, the small graphite lamellas are completely dissolved in the iron; this results in a state of high dispersion of the graphite in the castings, which in turn produces a fine grained structure on cooling. If coarse-grained iron is used, the graphite lamellas are not uniformly distributed, this leads to irregularities in structure on cooling. A. PAPINBAU-COUTURE

Sulfur in cast iron. H. H. EBERT. *Zentral Europ. Gießerei-Ztg* 3, 3-4(1930).—Methods of desulfurization are discussed. B. C. A.

Cast chromium steels in wider use. H. D. PHILLIPS. *Steel* 88, No 2, 39-43(1931).—A discussion of production processes, properties and uses of chromium steels. LESLIE D. DRAG

Overcoming certain operating troubles in making open-hearth steel. T. N. ARMSTRONG. *Iron Age* 127, 864-7(1931).—The advantages incident to the use of burnt lime instead of raw limestone in basic open hearth steel charges are discussed, as well as the methods of overcoming a no of common troubles encountered in open-hearth

operation, particularly those connected with the use of Southern grades of pig Fe.

Downs Schlaaf

The system. iron-carbon. Eugène L. Duruy *Rev metal* 27, 680-92(1930).—

II. S. V. Kloostera

A review

Manufacture of ferro-silicon. Roy P. Hudson *Blast Furnace Steel Plant* 19, 109-101(1931)—The chief district for producing ferro-Si is Jackson County, Ohio. Originally, lean ores of the district were used, but these have been supplanted by Lake Superior ores. The blast furnace can produce metal contg 18% Si, but most grades average 6-15% Si. Ferro-Si is notably low in S on account of the high temp of formation and large vol of slag produced. Bituminous coal is used in admixt. with coke to decrease the fuel cost. Because of the high temp required to reduce a large quantity of silica to Si, the heat should be coned in a small area, therefore, the bosches and hearths are always smaller. In the manuf of high Si iron the evidence indicates that when alumina is lower than 17% it acts as an acid and above this it acts as a base. But as it rarely exceeds that amt., alumina is always considered an acid. A. W. Holmes

Open-hearth combustion. W. P. Chandler, Jr. *Proc. Eng. Soc. West Penn* 46, 241-63(1930)—See *C. A.* 24, 5265

II. C. Duvv

Gas and stock flow in the blast furnace. S. P. Kinney. *Blast Furnace Steel Plant* 19, 407-11(1931)—See *C. A.* 24, 808.

A. W. Holmes

Recent advances in metallography. I. L. Houghton. *1st Communications New Intern Assoc for Testing of Materials A*, 316-9(1930)—A review of the advances in the methods of thermal and microscopic analysis of alloys, and in the use of phys property measurements as an aid to the study of constitution. Attention is also directed to some of the other principal problems such as age hardening, modification and the problems raised by allotropic modification, intermetallic compds, etc. E. J. C.

Bearing of present-day knowledge of plastic deformation on the testing of metals. G. Sachs. *1st Communications New Intern Assoc for Testing of Materials A*, 309-15 (1930)—In the testing of metals and other cryst. materials it will in the future be necessary to bear in mind the results of modern investigations on the following points:

(1) The properties of a material may be different in different directions. (2) The values of the elastic consts may be seriously affected by previous treatment and by the methods of measurement. (3) The elastic limit and the yield point are easily affected by previous cold working. (4) The previous history of a material can subsequently be detd from the distribution of internal stresses within it. (5) In the calcn of structural parts it is necessary to take account of the new knowledge concerning the liability to deformation under normal and shear stresses. (6) The behavior of materials when subjected to technological processes can be calcd on the basis of their properties as detd by testing methods and can be checked by specially designed methods of testing. A bibliography is included. E. J. C.

Materials at high temperatures. General properties, limiting creep stress and limit of proportionality. R. G. Barton. *1st Communications New Intern Assoc for Testing of Materials A*, 74-9(1930)—The general behavior of steels at high temps under short time tensile tests and the effect of prolonged loading in producing creep or flow and intercryst. failure are discussed. Limiting creep stresses are described and the values obtained for them are compared with results from limit of proportionality detns and time yield tests as defined by Hatfield (*C. A.* 23, 4429). The safe stresses to be used for design purposes are stated. E. J. C.

Materials at high temperatures. A. Power. *1st Communications New Intern Assoc for Testing of Materials A*, 80-4(1930)—Tensile strength at high temps is materially affected by the rate of testing. The time effect, however, is materially less in the detn of the yield point or the 0.2% limit. Published results generally indicate that the yield point at high temps, for particular classes of material, bears a definite ratio to the tensile strength at room temp. In place of the lengthy tests under prolonged loading, Pöppel and Dahmen propose an abridged method of testing which in the course of 1 or 2 working days permits of an approx detn of the so-called limiting creep stress. This is regarded as being attained when the rate of extension in the 3rd and 6th hrs. after applying the load does not exceed 0.001%. A bibliography is given. E. J. C.

Behavior of metals at high temperatures. H. Stäger. *1st Communications New Intern Assoc for Testing of Materials A*, 89-90(1930)—Tests to det the limiting creep stress were made at 400°, 500° and 600° on steel castings of widely varying compns. and on a no. of alloy steels. Rapid high-temp tests afforded no indication of the value of the limiting creep stress. For the design and construction of machine parts intended to work at high temps only the actual creep stress can be employed. An exact defi-

nition of this stress can only be arrived at as the result of further investigation. It was found that it is not the steels contg large amts of alloying elements which give the best strength results at high temps, but rather those in which the alloying elements have been added in the most suitable amts for the purpose of strengthening the material lying in the crystal boundaries. E. J. C.

Materials at high temperatures. Special alloys of iron, nickel and chromium. ROBERT HADFIELD. *1st Communications New Intern Assoc for Testing of Materials A*, 104-10(1930).—A review of the development of heat resisting alloys and their use in the improvement of furnace efficiencies as metallic recuperators, conveyors for the mech operation of furnaces, exhaust valves for internal-combustion engines, steam turbine blading and steam fittings. Of these alloys, the Fe-Ni-Cr type are probably the most useful and important. Most of the practical alloys of this type fall into 3 groups, 1 contg iron, say, 18-25% of Cr, with 7-10% of Ni, the Fe content amounting to 65-70%. The second range is higher in Ni its percentage being in this case from 20 to 40, the Cr amounting to 10-15% and Fe 60-60%. The third group includes alloys contg about 65% of Ni, 12% of Cr and 25% of Fe. Other elements in smaller percentages are added in many instances. E. J. C.

The micromechanical study of metals. P. A. WELIKOV, N. P. SICHIAPOV AND W. P. LORENZ. *1st Communications New Intern Assoc for Testing of Materials A*, 328-38(1930).—The micromech method used by the Inst Sci 1 xptl des Transports, Moscow, for studying the mechanism of plastic deformations occurring in the materials employed in means of transport, comprises a combination of the micrographic investigation of the deformations with mech tests on specimens of small dimensions, the 2 being carried out simultaneously on the same test piece. Tests were carried out on compressive and tensile strength, and in the latter case test specimens both with and without a notch were used. The following materials were investigated: mild steel used in bridge construction, steel for rails, fine-grained Cr and Ni steels. The rept describes the development of permanent deformations observed. E. J. C.

Nonmetallic inclusions in metals. G. P. CONSTOCK. *1st Communications New Intern Assoc for Testing of Materials A*, 348-51(1930).—The causes, effects, identification and elimination of nonmetallic inclusions in steel and nonferrous alloys are discussed briefly. Illustrations of the most common type of inclusions are presented, as well as Wohlrman's scheme for their identification by metallographic means. Proper polishing of specimens is important. The various types should be considered separately, since the causes of their existence in metals, and the means for their elimination, are different for the different types. The effects of their presence are usually, but not always, undesirable, and the prevention of segregation is of more general importance than the elimination of inclusions. E. J. C.

Physical and chemical properties of the light metals. Prescription for the mechanical testing of aluminum alloys. A. VON ZEELEON. *1st Communications New Intern Assoc for Testing of Materials A*, 257-72(1930).—The phys consts of Al and Mg are given and their alloys described. The standard specifications for tests on casting and forging Al alloys are stated, as they are recommended by the Swiss section of the New Intern Assoc for the Testing of Materials and by the European Al producers. They concern the detn of tensile strength, elongation, limit of elasticity, upper yield point, modulus of elasticity, Brinell hardness, resilience, bending capability, casting property and contraction for casting alloys and construction for forging alloys. E. J. C.

Notched-bar impact test—standard test piece. M. MOSNA. *1st Communications New Intern Assoc for Testing of Materials A*, 223-4(1930).—The Charpy test piece is not altogether satisfactory. Both producers and users of materials are looking for a new test piece on the following lines: (1) The specimen must be small so as to enable each part to be tested without unreasonable expense. (2) The shape must be such that the various mech conditions of the material are widely differentiated at room temp. (3) The range of application must be as wide as possible so that auxiliary test pieces are only necessary in exceptional cases. (4) It must be easy to make, the tools commonly available in a workshop are used. Various proposals are illustrated. The German standard notched bar test specimen, the principal features of which are already fixed, corresponds closely to the test specimens used in other countries. E. J. C.

Notched-bar impact tests—standard test piece. R. PAWLISKA AND M. SCHMIDT. *1st Communications New Intern Assoc for Testing of Materials A*, 217-21(1930).—No simple proportionality exists between notched bar test specimens of various forms, and the test results are influenced by the temp, shape of the specimen and by the speed of the test. For plain and alloy steels, distinction must be made between high and low values for the impact toughness, the development of these, with reference to the

the 3rd and 6th hr. In alteration of Pömp and Dalmón's previous suggestions the creep limit of straight C steels is defined as that max. load at which the velocity of extension does not exceed 0.0001% between the 5th and 10th hr. G. G. NEUMARK.

Prospects hopeful on metals for 1000-1500°F. H. W. SKEVIN. *See World* 97, 459-61(1931). Greater demands are being made for steel to withstand high temps. under high-pressure conditions. Above the strain hardening or so called equal cohesion temp. there is no such thing as "creep limit," creep constantly going on, but more rapidly with higher temps. and increased stresses. Large grain sized material has greater resistance to creep than the same material with reduced grain size. Fatigue of over-stressed metal is much more important at lower temps. than at high operating temps. Boiler shells, headers, piping, etc. have a certain flexibility, but valve seats require new and more serviceable materials than the common C steels. Brass, Monel metal, Ni and Ni alloys, stainless steels and some harder alloys are used, depending upon the service desired. A proper valve seat should be noncorrosive, have a proper coeff. of expansion, have forging or casting quality, machinability, density and smooth texture, proper strength and ductility, hardness to withstand scratching, nongalling or seizing quality and be reasonable in cost. "Astralloy" is being used for the purpose.

W. H. BOYNTON.

Advances in microscopy. F. F. LUCEY. *1st Communications New Intern. Assoc. for Testing of Materials A*, 3rd S.(1930) of C. A. 23, 73. Specifications are given for new metallographic equipment expected to yield satisfactory results at very high magnifications (5000-7500 diam.). Developments in progress in the field of ultra-violet metallography are discussed. New optical glasses of ultra-violet transmitting characteristics are needed and the possibilities are reviewed briefly. The advances to be expected in metallography, especially with respect to the phenomena involved in the hardening of steel and the nature of the constituents, martensite and troostite, are discussed.

R. J. C.

Progress of microscopy. H. WATKINSON. *1st Communications New Intern. Assoc. for Testing of Materials A*, 3rd S.(1930) of C. A. 23, 73. Progress in metallurgical microscopy in Great Britain since 1920 is described. An outline is given of the work of the Roy. Microscopical Soc. in the interests of technical microscopy, of new developments in app. and of advances in microscopical and photomicrographic technique. Reference is made to the work in progress at the Imp. Coll. of Sci., where the technique of ultra-violet light microscopy is being developed.

R. J. C.

Strengthening by reversals of stress. W. SCHWENNING AND R. STROBEL. *Z. Metallkunde* 22, 162-4(1930).—Conclusion of a previously published article (C. A. 23, 1107). This article gives an account of the changes in structure (slip lines, grain boundaries, etc.) and the effect of heating (recrystallization) upon the structural changes observed.

ROBERT F. MEHL.

Theory of solution of metals. H. M. STRAUMANN. *Z. physik. Chem., Abt. A*, 153, 107-11(1931); of C. A. 24, 4688.—Polenikual with Thiel (C. A. 25, 973).

OSCAR T. QUIMBY.

Note on the preceding contribution [on theory of solution of metals]. A. THIEL. *Z. physik. Chem., Abt. A*, 153, 112(1931).—Polenikual with Straumann (cf. preceding abstr.).

OSCAR T. QUIMBY.

Thermal analysis and dilatometry. H. SCOTT. *1st Communications New Intern. Assoc. for Testing of Materials A*, 3rd S.(1930) of C. A. 23, 41. —The salient features of methods for revealing physicochem. changes in metals at high temps. have been critically examined, thermal analysis, dilatometry, resistance testing and magnetic testing alone are considered. Thermal analysis alone is well adapted to the detection of solid to liquid phase changes. Its chief shortcomings are inability to follow changes at a const. temp. and to measure quantitatively a sp. phys. property. The other test methods are complementary to thermal analysis in these respects. Dilatometry has the distinction of revealing under favorable conditions all types of discontinuous changes in solid metals. The resistance measurements are particularly adapted to the detection of changes in solid only. Magnetization tests are superior for the detection of phase relations in certain steels and permit the identification of sp. magnetic compounds in steel. There are several types of equipment available for each test method. Of those for thermal analysis, the Rosenhain furnace with the inverse rate method is the most flexible one. An autographic device for plotting such curves has been developed. The fully automatic dilatometer of Chevenard is very convenient for detecting physicochem. changes and furnishes fair values of expansivity as well. Elec. resistance and magnetization tests have unique applications, but require further development to become generally acceptable lab. tools.

R. J. C.

Heat treatment of aircraft parts. HORACE C KNEER *Proc Am Soc. Testing Materials* 30, Pt II, 151-70(1930).—A general discussion. The properties of various SAE steels under different heat treatments are tabulated. For Cr Mo aircraft steel 4130X, Kneer advocates 1575-1625°F for normalizing or quenching, and 1275-1375°F. air or 1375-1425°F furnace, for annealing. On oil quenching, and tempering at 680-1100 F yield points from 180,000 to 110,000 lb/sq in are obtained. Types of furnace equipment for various heat treatment operations are discussed. H. W. G.

Heat treatment of carburized parts. H. M. McQUAIN *Iron and Steel of Canada* 14, 22-4(1931).—A considerable amt of steel, which at the present time is being given the costly double treatment subsequent to carburizing, might quite successfully be quenched directly from the carburizing box. The 3 methods of single-quenching carburized work are discussed, and the types of steel suitable for use when these methods are employed are described. A statement is given of the proper procedure to follow in establishing the proper heat treatment of a carburized part based on the assumption that the analysis selected has been chosen as the result of a careful study of the stresses and service to be met with in use. O. W. FELLIS.

Heat treating pipe-line couplings. C. B. PHILLIPS. *Steel* 87, No 26 43-5(1930).—The plant of the Dresser Mfg Co at Bradford, Pa., is described. L. ASLIE B. BRAGO.

Red shortness. FRANZ HATLANEK *Arch Eisenhüttenw* 4, 207-14(1930).—In addn to red shortness due to oxides, brittleness is also due to other causes at high temps., among them the action of gases and the presence of Mn. Brittleness produced by O at relatively low temps. can be prevented by the presence of C, Mn, Si and Al. Blue shortness is caused by oxide inclusions, is produced earlier than red shortness, and is more difficult to overcome. H. STOEZT.

The measurement of the true specific heat of solid and liquid metals at high temperatures. H. SEEMANN *Z anorg allgem Chem* 195, 345-65(1931).—S describes a process for the measurement of the true sp. heat in which a W spiral is employed as the source of energy. Complete details of the method of supporting the spiral within the sample of metal to be investigated are given. Four diagrams are given which describe the app. used and the elec. equipment employed in connection with the expts. Particulars are given of the results of expts. on Cu, Al and Mg, the true sp. heats of which were measured in the solid state. Curves are given, showing the relationship between S's results and those of previous workers in this field. The application of this method to the investigation of liquid metals is described. In this case, the W spiral is protected by means of a porcelain sleeve. Measurements of the true sp. heats of solid and of liquid Ti were made from the range 18° to 500°. It is shown that Ti undergoes a transformation at the point of 225.7°, the heat of transformation amounting to 98 cal/g-atom. O. W. FELLIS.

Standardization of methods for testing the fatigue resistance of metals. K. LAUBE *1st Communications New Intern Assoc for Testing of Materials A*, 111-8(1930).—The alternating stress resistance of a material is detd. by the endurance limit under a standardized system of dynamic loading. Its value is dependent on the 2 factors of fatigue and work hardening. Abridged methods of testing fail because work hardening is not completed even after a million reversals of loading. The various methods of stressing, either by steadily changing loading or by alternating tension and compression cannot be regarded as interchangeable. The fatigue range varies with the chem. compn. of the material. It is diminished by occasional over loading and depends very much on the smoothness of the surface of the test piece. Much less important are the effects of mech. or thermal treatment prior to testing, while the variations of temp. during the test itself and the frequency of alternation have no effect. The influence of the actual dimensions of the test piece is probably negligible. The great importance of the character of the surface in itself constitutes a very serious obstacle to standardization of the test. Further, the difficulty of ascertaining and keeping to a definite max. of loading in practical use makes the application of the fatigue limit for practical purposes somewhat illusory. F. J. C.

Fatigue of single crystals of pure metals. H. J. GOUGH *1st Communications New Intern Assoc for Testing of Materials A*, 133-44(1930), cf *C A* 23, 2402.—The characteristics of deformation, under fatigue stresses, of single crystals conforming to each of 3 crystallographic forms, i. e., face-centered cubic (Al), body-centered cubic (α Fe) and close packed hexagonal (Zn), are summarized. It is shown that the criterion of plastic deformation is, in every case, the max. value (max. resolved shear stress) of the shear stress component on the slip plane along one of the slip directions contained by the slip plane. The direction of slip is always parallel to the "principle line" of

atoms (that having a greater no. of atoms per unit distance than any other line of atoms in the lattice) In the face centered and close packed hexagonal systems, the slip plane is the plane of max. slip. α Fe deforms by a process of duplex slip on planes of high slip. Fatigue cracks are initiated in the region of max. resolved shear stress, not affected, apparently, by normal stress considerations. The propagation of the crack is, however, largely assisted by tensile normal stress. The direction of propagation is often dictated by max. not max. resolved, shear stress. Reference is also made to certain theoretical deductions arising from the experiments. A bibliography is included. E. J. C.

The effect of surface condition on fatigue test results. C. H. M. JENKINS. *1st Communications New Intern. Assoc. for Testing of Materials A*, 145-8 (1930).—An attempt has been made to translate fatigue tests to the service conditions of such machine parts as springs, shafts with keyways, etc. The effects on the test results of such factors as the form of section, the surface finish and abnormalities in composition and micro-structure are discussed. The accelerating influence of chem. corrosion is marked, and perhaps phys. abrasion of the surface may also weaken the material. Although the general tendency of imperfect surface conditions is to reduce the value of the fatigue strength, there appears to be a slightly increased resistance to fatigue in the case of certain surface hardened materials. A bibliography is given. E. J. C.

Fatigue, I. Relation of the fatigue limit to the elastic limit and other mechanical properties. II. Effect of surface condition on fatigue test results. P. LUDWIG. *1st Communications New Intern. Assoc. for Testing of Materials A*, 119-32 (1930).

Fatigue studies of telephone cable sheath alloys. J. R. TOWNSEND AND C. H. GREENALL. *Proc. Am. Soc. Testing Materials* 30, Pt. II, 393-405 (1930).—In the study of fatigue properties of cable sheath alloys, a preliminary survey is made by use of a const. deflection method on a flat cantilever specimen at a single deflection, under which pure Pb fails in 0.3 to 1.3 million cycles and a 1% Sb alloy in 1.8 to 2.0. Alloys of refined Pb with 0.04% Ca ran for 13 million as a min. up to 169 million unbroken. Very small quantities of Sn or Sb injure the Ca-Pb alloy on the test and it is stated that it will detect the effect of 0.001% As, which is not shown by tensile test. Actual endurance tests are made on a cantilever type rotating beam machine. Refined Pb, unalloyed, shows a 50 million cycle endurance limit at 215 lb./sq. in. The same with 1% Sb shows 45 and with 0.04% Ca 850 lb./sq. in. S-N curves are given. Static flow tests show an elongation in 3000 hrs. at 920 lb./sq. in. of 0.02" in 2" for the Ca alloy and 0.10" for the 1% Sb alloy. For metallographic and other data on these alloys, see Schumacher and Bouton (*C. A.* 24, 2109). H. W. GILLET.

Constancy of the work done to cause fracture as an explanation of fracture by fatigue and other loads. K. LJUNGBERG. *1st Communications New Intern. Assoc. for Testing of Materials A*, 149-54 (1930).

Relations between the changes in physical properties by cold working and mixed crystal formation. CURT AGTE AND KARL BECKER. *Physik Z.* 32, 65-80 (1931).—Cold working and the introduction of impurities cause similar changes in many of the physical properties of metals. A review is given of data on W showing these changes. Bibliography includes 75 references. T. H. CHILTON.

The theory of recrystallization. J. A. M. VAN LIEMPT. *Z. anorg. allgem. Chem.* 195, 360-86 (1931).—Having assumed (1) that in the solid state only those atoms can exchange places, which possess an energy equal to, or greater than, a given value $E = ar_0^2$, where r_0 is the min. amplitude required to cause a change of place, and (2) that amplitudes in excess of r_0 are impossible, and having accepted Lindemann's assumption that at the m.p. of a solid the amplitude of vibration of the atoms is approx. equal to the mean distance of the atoms apart, van L. develops a general diffusion formula applicable to single crystals, from which, if the distance between the atoms of the metal in which diffusion occurs, the characteristic vibration frequency, and the m.p. of the diffusing metal, and the temp. at which diffusion occurs are known, a value for the diffusion const. at the above temp. can be obtained. Hence, the relationship between the diffusion const. and the temp. of the metals can be calculated. A relationship between the diffusion const. of deformed and undeformed metals is then evolved, and the fact that the rate of diffusion of a deformed metal in an undeformed metal is greater than that of the undeformed metal is explained. Modifications necessary to apply the theory to polycryst. materials are discussed. van L. then derives general formulas for the relationships between temps. of annealing and the times required to effect complete softening at these temps., and between the m.p.s. and the temps. of recryst. of the metals. Tables showing the relationship between observed and calculated values are given.

The phenomena of recrystallization in solid solutions and of recrystallization in general are discussed.

O. W. FILLIS

The aeration theory of Evans. E. HERZOG AND G. CHAUDRON. *Korrosion Metall-schutz* 7, 1(1931).—A reply is made to objections by Maass and Liebrich (*C. A.* 24, 3744, 5705) against conclusions drawn from 'drop' experiments. Details of these experiments are given, which uphold the conclusions of the authors and the aeration theory by Evans. E. MAAS AND F. LIEBRICH. *Ibid.* 2, 3.—The unetched spot in the center of a drop observed by Herzog and Chaudron (above) after adding to this drop a NaCl solution containing 1% H_2O_2 is assumed by the authors not to be due to a different aeration, but to be due to disturbances of convection currents caused by the introduction of the capillary pipet. It is important whether the pipet only touches the surface of the drop, or whether it penetrates beneath it. In the first case, the solution introduced will not go down to the center of the drop, as assumed, but will, provided its density is approximately the same as that of the drop, diffuse slowly throughout the drop. In the second case, a disturbance of the convection currents takes place, which will cause the appearance of an unetched or only slightly etched spot in the center, regardless of whether the solution contains H_2O_2 or not. A similar unetched spot can also be obtained if, in place of a pipet, a thin glass rod with a plane polished end is immersed for only 5 or 10 min. into the drop. This cathodic spot formed in the center remains unetched even after the removal of the glass rod.

LEONOLD PRESSEL

Investigation on the behavior of metal toward cleaning and disinfecting materials. M. I. SCHULTZ. *Maketen Ztg. (Hildesheim)* 44, 1439-41, 1461-3, 1495-1501, 1515-6, 1557-60(1930).—Metal sheets of 100 sq. cm. surface are immersed for 24 to 48 hrs. at 45° to 50° in a solution of the material. Change in the metal, loss in wt., tempering color, surface and pore corrosion are noted. The results are placed in a table so that the proper material can be selected for each metal. There are no cleaning solutions which do not attack metal, but some are worse than others. The protection given by the addition of water glass and chromate is studied. The addition of chromate to the solution is limited by the cost. The amount of water glass to add to solutions of a number of common cleaners is given.

GEORGE R. GREENBANK

Metallurgy of the modern die set. JOHN H. HAVKA. *Iron Age* 127, 702-5(1931).—Castings for die sets are made almost exclusively from semi-steel or a special grade of gray cast Fe. The heat treatment of die set castings is described. Increased temperature decreases hardness, prolonged heating enlarges graphite flakes of matrix and decreases transverse strength. Leader pins and bushings are made of pearlitic Mn steels. Box carburizing is desirable in long leader pins and bushings. The rotary furnace has certain outstanding features so far as quick production and low cost are concerned. Drawing may be done in rotary furnaces, salt pots or oil tanks with accurate temperature regulation. Results of analyses are tabulated.

V. S. POLANSKY

Atomic hydrogen occluded in iron nitride. SHUNICHI SATOH. *Bull. Chem. Soc. Japan* 5, 291-303(1930).—The presence of atomic H in Fe nitride is confirmed. The evidences are: the liberation of H when Fe nitride is heated above 430° and decomposes slowly; the measurement of the single potential of Fe nitride in the normal FeSO₄ solution; the oxidation of atomic H by O in the water in which Fe nitride is immersed; the transformation of K ferrioxamide. Nitriding changes compact Fe into porous Fe. Light does not produce any marked photochemical effect upon the single potential of Fe nitride.

ALBERT L. HENNE

Nickel in aviation. GENERAL GRARD. *Aircraft Specialist* 5, 566-73(1930).—The properties of many steels are given.

A. J. MONACK

Discontinuous changes in length accompanying the Barkhausen effect in nickel. C. W. HEATON AND A. B. BRYAN. *Phys. Rev.* 36, 326-32(1930).

BERNARD LEWIS

Silver and its application to chemical plant. DONALD McDONALD. *Soc. Chem. Ind., Chem. Eng. Group* Feb. 13, 1931, 12 pp.; *Chemistry & Industry* 50, 168-78(1931).—The properties of Ag are reviewed with regard to its use in the construction of chemical equipment. The economic history, occurrence and metallurgy of Ag are discussed at length. The physical and chemical properties reported in the literature are critically reviewed with some additional observations. The power of absorbing O by Ag in the molten state has been studied in detail. The more important alloys of Ag and their applications are mentioned. The properties of Ag which make it attractive for the chemical plant are: its freedom from oxidation, resistance to acids, high thermal and electrical conductivities, and excellent mechanical characteristics. Various uses in the industries are mentioned. The high salvage value of Ag apparatus makes its use as a construction material more attractive.

ALLEN S. SMITH

The resistance of chromium-plated plug gages to wear. HARRY K. HERSCHMAN. *Bur. Standards J Research* 6, 295-304(1931) R. RIMBACH

The structure and the mechanical properties of pearlite. N. T. BELVAEV. *Rev metal* 27, 680-5(1930)—B indicates the possibility of using the interlamellar distance in pearlite as a measure for the mech properties of eutectoid steels (cf C A 23, 5454).

H. S. V. KLOOSTER

The effect of certain elements upon the segregation of cementite and its relation to the modification process. J. H. ANDREW. *Roy Tech Coll Met Club J* (Glasgow) 1929-30, No. 7, 16-7—Speculation upon the relation between the effect of 0.25% Mo in a 3% Ni steel, or of 0.60% Mn in a 0.23% C steel, in which the Mo or Mn is said to hinder free movement of carbides, preventing them from coalescing and thus producing a fine sorbite, and that of Na in the modification of Al-Si alloys H. W. G.

Thermochemical data on iron. W. A. ROTH, H. UMBACH AND P. CHALL. *Arch. Eisenhüttenw* 4, 87-93(1930), cf C A 24, 2039—The authors used azeotropic 20% HCl soln at temps in the neighborhood of 100° as a solvent for various Fe compds and detd the following heat effects, calcd at 19° $\text{Fe} + 0.5\text{O} = \text{FeO} + 64.2 \text{ Cal}$, $\text{FeO} + \text{Fe}_2\text{O}_3 = \text{Fe}_3\text{O}_4 + 7.0 \text{ Cal}$, $2\text{FeO} + \text{SiO}_2 \text{ (quartz)} = \text{Fe}_2\text{SiO}_4 + 8.1 \text{ Cal}$ and $\text{Fe}_2\text{O}_3 + 3\text{Cl}_2 = 2\text{FeCl}_3 \text{ (cryst)} + 1.5\text{O}_2 + 7.4 \text{ Cal}$ H. S. V. KLOOSTER

Alterations in the properties of hematite pig iron cast from the first melting. E. PIWOWARSKY. *Stahl u Eisen* 50, 966-8(1930)—The alterations which take place in the mech properties of pig Fe during tapping from the blast furnace into an acid mixer from which finished castings are poured have been detd. The compn. and properties of samples cast in sand and in chill molds directly from the blast furnace during the charging of the mixer and during the emptying of the mixer are tabulated, the variations over a period of 7 days being shown. Sand castings have the better properties when taken directly from the blast furnace, but chill castings are superior to sand castings poured from the mixer. In both cases an appreciable improvement in properties takes place during the passage of the metal through the mixer B. C. A.

Cast iron. E. PIWOWARSKY. *1st Communications New Intern. Assoc. for Testing of Materials A*, 5-9(1930)—The special properties of cast Fe (lack of quasi isotropy and influence of graphite inclusions) have for a long time made it difficult to introduce standardized methods of testing. The absence of reliable inter relationships between the mech. and phys properties necessitates furthermore a large number of sep. investigations to det. the quality. The following measurements are now made: (1) *The static properties*—(a) Tensile strength. Tensile tests are provided for. The fact that cast Fe is scarcely ever subjected to pure tension in practice makes this test of little value, but it is easy to carry out and to evaluate, (b) bending strength. The deflections must be given for as many loads as possible, in any case for the max load at fracture; (c) the Brinell hardness, (d) compression strength, with simultaneous recording of the stress strain diagram and the load at which cracks begin to appear; (e) shearing strength. The test for shearing strength does not give any simple relation between shearing and tensile strength. (2) *The dynamic properties*—The detn of the fatigue limits will continue to be of value until smaller foundries are able to procure rotating type fatigue testing machines. (3) *The elec and magnetic properties*—Experiences in testing Fe and steel may be applied. (4) *The machinability*—The hardness boring test of Kcep, Loewe or Kessner is widely used. Nevertheless the test on turning time to det. the cutting speed seems to be the only suitable one. The Brinell hardness for cast Fe bears no fixed relation either to the machinability or to the resistance to wear. (5) *The volumetric stability*—This test still lacks any degree of uniformity. (6) *The sp. vol*—This test is best carried out with the method of buoyancy. Measurement of the perviousness to gases appears to be of less importance. (7) *The ability to fill molds*—Rapid and practically useful results are given only by pouring tests. The tendency to porosity and to the formation of blow-holes is shown by Cook's K-test. E. J. C.

Cast iron—methods of testing. E. DÜBI. *1st Communications New Intern. Assoc. for Testing of Materials A*, 10-24(1930)—A separately cast test piece of 50 mm diam. gives results which agree very closely with those obtained from test pieces of 30 and 50 mm. diam. machined from the main casting. Small test pieces taken according to the French specification from the main casting give valuable criteria of the nature and uniformity of the internal structure of the casting itself. As regards the strength of the material, however, these small test pieces do not give any more accurate indications than specially cast test pieces having a diam. corresponding as nearly as possible to the thickness of the wall of the casting under investigation. The strength properties of casting apparently can only be detd. conclusively from tests on the casting itself. E. J. C.

High-duty cast iron. GEDFORD SUTHER. *1st Communications New Intern Assoc for Testing of Materials A*, 49-63(1930)—The methods of producing high grade cast Fe are discussed with particular reference to the effect of a considerable reduction of C content on the properties of the Fe. As the C content diminishes, bending strength, tensile strength and resistance to impact increase, while in general the deflection remains unchanged. The shrinkage and the tendency to form cavities, however, increase, as do the general difficulties of foundry technique. The pearlitic ground mass of high-grade cast Fe has in the cast condition, a tensile strength of approx. 70 kg per sq mm. By means of alloying addns., so long as the structure remains pearlitic, the Brinell hardness of cast Fe is increased while the other strength properties are not materially changed. Carbide forming elements such as Cr, W, Mo and V increase the resistance to abrasion and to thermal effects, but reduce the machinability. E J C

The results of investigations on the growth of cast iron. FRANZ ROLL. *Die Gießerei* 17, 995-8(1930)—The views of various investigators on the influence of the compn and other factors on the growth of cast Fe are tabulated. J BALOZIAN

Cast-iron testing in Great Britain. J G FRASER. *1st Communications New Intern Assoc for Testing of Materials A*, 1-4(1930)—A specification which has the approval of the founding industry was issued in 1928 by the Brit. Eng. Standards Assoc. While it covers only 2 qualities of Fe for general engineering castings, the testing procedure is applicable to all types and qualities. The features of this specification are as follows. The size of the test bar varies with the thickness of the casting, 3 sizes being used at present. The test bars are cylindrical and may be cast away from the casting. Both tensile and transverse tests are covered and no chem. analysis is demanded. The majority of engineering castings will be covered by the intermediate size of bar 1.2" in diam which is virtually the standard bar used on the Continent and in America. The adoption of this basis for testing has stimulated a considerable amt. of investigation on the relations between mech. properties, but it is believed that at the present stage of development no other mech. tests can be usefully employed in general specifications. E J C

Heat- and acid-resisting cast irons with high chromium and carbon contents. EMANUEL VALENTA. *Iron and Steel Inst (London) Carnegie Scholarship Mem* 19, 79-165(1930)—Three series of Cr-Fe C alloys were studied. (I) without added elements, (II) with addns. of Si and (III) with addns. of Al. (I) In this series the C content varied from 0.15 to 3.6%, while the amt. of Cr was between 4.7 and 36.0%. The Si content was 0.4% for low Cr and 0.8% for high Cr percentages. It was found that above 16% Cr the resistance to scaling is perfect. Beyond 35% Cr the heat resistance does not increase further. The proportion of primary double carbides increases with the C content. These carbides form long, heat resistant needles which destroy the homogeneity and strength of the castings. Hence the hypereutectic alloys are of no technical importance. Below 0.5% C the alloys are more expensive and must be produced in an elec. furnace. Test bars, both prismatic and cylindrical, were prep'd for chem., metallographic, dilatometric, thermal and mech. examn. Brinell hardness detns. were made after casting, annealing and air cooling for various thicknesses and different rates of cooling. Graphs are presented showing lines of equal hardness depending on the Cr and C contents. For machinability, if a hardness of 350 is assumed to be the limit, it is necessary to keep the C content below 1.6 or 2% in alloys contg. 25 or 35% Cr, resp., in annealed specimens. For high resistance to wear and abrasion a Cr/C ratio of 10/1 may be used with advantage. The useful alloys in the Fe-Cr-C system are those contg. 1.8-2% C which may be divided in 3 zones: (a) alloys in which the pure δ -crystal predominates, strongly corrosion resistant, suitable for forging and without hardening ability, (b) alloys showing both γ - and δ -crystals with reduced corrosion resistance and lower malleability and (c) alloys of pure γ -crystals with materially reduced resistance to corrosion, recommended for service at high temps. (II) The examn. of alloys contg. Si in amts. varying from 0.5 to 4.3% with Cr ranging from 20 to 35% showed that the eutectic valley is displaced to lower C percentages. Hence the range of useful hypoeutectic alloys is much lower. Addn. of 1% Si displaces the eutectic line to lower C content by about 0.3% C. The corrosion resistance is greater than for those alloys contg. no Si. With constant Cr content the effect of varying amts. of Si (0.5-2.3%) is to raise the temp. of entrance into the γ -phase. For 1% Si the rise is about 80°. The effect is less for lower C content where the eutectoid temp. rises rapidly. Hence the lower the C and the higher the Cr percentage, the smaller is the amt. of Si required to suppress the hardening of alloys cooled from 1000°. For high C alloys as much as 4% Si might be required to prevent air hardening, but the product would be brittle and unmachinable. (III) Alloys with Al as addn.

element contained 1.1-2.0% C, 10.3-29.8% Cr and 0.6-2.4% Al. Even a low percentage (0.5%) of Al was in most cases sufficient to suppress entirely the self-hardening properties. The effect of Al in the raising of the critical points and upon the structure was about twice that of Si. Air cooling gave alloys with low hardness, shock resistance and good malleability. The location of the eutectic line was not materially altered by adding 1-1.5% Al, hence the range of useful high C alloys is greater than that for Si alloys. The alloys studied are classified on the basis of the dilatometric curves obtained. Numerous tables, charts and photomicrographs are given. A suitable reagent for etching these alloys consisted of a cold a/c. soln. of 5% picric acid, 2.5% HCl and 2% FeCl₃.

H. S. V. KLOOSTER.

Nickel-vanadium and nickel-molybdenum cast irons. II. J. RAN CHALLANSONNET. *Rev. Metal.* 27, 634-71 (1930), cf. *J. A.* 25, 1197. —The dilatometric and microscopic study of a synthetic cast iron contg. 3.7% C showed that the presence of 0.2% and 0.5% Mo either alone or with 1 and 2% Ni, resp., does not affect the Curie point of Fe₃C. Addition of 0.2% and 0.5% Mo raised the Ac point 25° to 725° and in alloys with 1 and 2% Ni the added Mo raises the Ac point from 660° and 665° to 725°. On cooling the Ar point is lowered and the effect is cumulative when both Mo and Ni are present. Reheating of the cast alloys brings about graphitization in the specimens contg. 1% Ni and 0.25% Mo (at 1000°), 2% Ni and 0.25% Mo (at 1050°) and 2% Ni + 0.5% Mo (at 1070°). The microstructure of all 6 alloys when fast cooled showed no graphite. Mo refines the pearlitic constituent and produces small bright spots, possibly complex Fe-Mo carbides which diminish in size in the presence of Ni. The com. cast irons were the same as those used previously for prep. Ni-V cast irons and distinguished as A, B and C, resp. In each group 6 alloys were studied contg. approx. 1% Ni + 0.25% Mo, 1% Ni + 0.5% Mo, 2% Ni + 0.25% Mo, 2% Ni + 0.5% Mo, 0.25% Mo and 0.5% Mo. All alloys as obtained after casting contained both graphite and cementite. Mo favors somewhat the cementite formation at low C contents (2.5%) but the graphitizing effect of Ni is only slightly changed by Mo. Dilatometric curves show sudden expansion on heating at temps. below the Ar point for all alloys contg. Ni. This expansion, due to graphitization, is absent in alloys contg. only Mo added when heated to 1000°. A study of the effect of the rate of cooling on the microstructure indicated that no complex carbides were produced in the com. cast irons but that Mo, like Ni, increases the hardness of the pearlitic matrix. With a C content as high as 3.5% C, Mo does not noticeably increase the percentage of combined C. Contrary to the effect of V on rapidly cooled cast irons contg. Ni, Mo does not prevent a martensitic quench. Tables, graphs and photomicrographs accompany the article. No definite conclusions could be derived from the data obtained.

H. S. VAN KLOOSTER.

Non-metallic inclusions in iron and steel. A. BAIKOV. *1st Communications New Intern. Assoc. for Testing of Materials A*, 355-62 (1930). —The non-metallic inclusions in cast metal are generally present in only very small proportions (tenths or hundredths of 1%), and usually consist of S and O compounds. The sulfurous inclusions are solid solns., and the O compounds are fused matter resembling a slag. If this view of the nature of inclusions is correct, then the way in which they are distributed is determined by the circumstance that the complete constitutional diagram of the system cast metal + material of the non-metallic inclusions exhibits the phenomena of limited solubility in the liquid state and forms 2 liquid layers. Non-metallic inclusions with a m.p. lower than that of the metal are of a totally different character. In this case the inclusions are in the form of drops and retain their spherical shape while surrounded by the solidified metal. The third type of non-metallic inclusions is observed in metals with low Mn content, considerable quantities of S and low C content. In this case practically the whole of the S forms Fe sulfide. With excess metal, the non-metallic inclusions solidify at the end of the crystal growth process, forming a skin of eutectic alloy around the grains of the segregated metal.

E. J. C.

Report of research committee on yield point of structural steel. W. O. M. WITHERY, *et al.* *Proc. Am. Soc. Testing Materials* 30, Pt. 1, 173-58 (1930). —In a cooperative study, the yield point of specimens cut from structural shapes of 10 different sizes and shapes, ranging from 1/4" plate to 26" girder beam and made from steel of 0.14 to 0.29% C, 0.37 to 0.73% Mn, was found to range from 29,000 to 39,000 lb./sq. in. depending on variations in rolling and cooling conditions, on finishing temp., on size of section and position of the specimen in the section. The ratio of yield point by drop of beam to tensile strength ranged from 0.50 to 0.65, hence it cannot be stated with sufficient accuracy from the tensile values. There was considerable discrepancy between mill tests and the more carefully conducted cooperative tests. A single specimen will not adequately represent the av. properties of heavy girder beams and H-column shapes.

and 3 specimens from intermediate flange, root and web should be taken. If only one is taken, it should come from the intermediate flange. The most representative single specimen is from an intermediate location in plates and angles and from the center of the web in channels and I-beams.

Mechanical properties of British steels. W. E. DALRY. *Proc. Inst. Civil Eng.* (London) 217, 25-61 (1928).—Stress-strains are taken from (1) the non load-to-break diagram, (2) the elastic diagram, (3) the push-pull diagram, and (4) the torque-twist diagram to indicate the stress at which slip begins for Bessemer and steel, open hearth basic, open hearth and (0.52% C), open hearth acid (0.62% C) and Ni-Cr steels. Elastic twist changes gradually to plastic twist, the change being more marked in Bessemer steel. The elastic diagrams are similar but limits of proportionality vary greatly.

The macroexamination of steel. Metallographic methods used at the Watertown Arsenal. M. G. YATSEVICH. *Army Ordnance* 11, 297-305 (1931).—Prep. methods are discussed and objects of tests, etc., are considered. Examples are given. Three tables on quantity of reagents, etching times, etc., are given. The interpretation of the picture revealed on the surface of the metal is discussed.

Selection and heat treatment of ball- and roller-bearing steels. ERNEST F. DAVIS. *Fuels & Furnaces* 9, 153-62 (1931).—Ball steels contain 1% C and 0.75-1.65% Cr, depending upon ball diam. Smaller sizes have lower Cr contents. Failures of balls in service result from surface defects, surface decarburization, excessive carbide segregation, and the commoner defects such as inclusions and center segregation. Steels for races are the high-C-Cr steel and the carburized grades such as the Ni-Mo, and the 3.5% Ni. Most widely used steel for carburized races is SAE 4115. Roll steels are the high-C-Cr steel, spring steel of modified SAE 4150 analysis and the carburized grades. Analyses are given of the various steels and methods of heat treatment are outlined.

The strength of steels at high temperatures. H. A. DIXON. *Roy. Tech. Coll. Min. Club J.* (Glasgow) 1929-30, No. 7, 33-7.—A brief general discussion.

Effect of rate of cooling on the structure and constitution of steel. J. M. ROBERTSON. *Safety in Mines Research Board Paper No. 59*, 57 pp (1930).—The effect of variation in the rate of cooling steel cannot be accounted for by the supposition that an increased cooling rate produces in turn white, troostite and martensite, as many other different forms of structure are evident between these states. It is suggested that the constitution of steel should always be considered in terms of the quantity of α -solid soln., γ -solid soln. and cementite present and their relation with one another. Four series of structures may be obtained by different methods of cooling. The first includes all structures consisting of ferrite and pearlite or the latter alone. Again, continuous rapid cooling which suppresses the Ar1 and lowers the Ar2 point produces the series of quenched steels, the structure of which may be varied considerably. When cooled so that the Ar1 point is suppressed and the rate then retarded so that the Ar2 change takes place at const. temp., a 3rd series results, and a 4th is similar except that the steel is withdrawn before the change is complete, the final change being effected by atm. cooling. The general arrangement of all ferrite-pearlite structures is determined by the manner in which ferrite develops from austenite, but the structure of the grains of pearlite is determined by simultaneously formed ferrite and cementite. The only structures which temper in a different manner are those which contain a considerable quantity of α -solid soln., since, during reheating, decomposition of this α -phase occurs by the gradual separation of C as the temp. is raised and not by a sudden precip. of fine C with subsequent coalescence to form larger globules.

Effect of the state of the cementite on the heat sensitivity, the tendency to carbide hardening and formation of hardness fractures in carbon steels. S. STRICKLAND. *Steel & Iron* 50, 1154-6 (1930).—The effect of various heat treatments on the microstructure and mech. properties of steels with 0.8 and 1.3% C has been investigated. The hypereutectoid steel after annealing at 900° and slow cooling becomes hard and tough after hardening at 725° and quenching. Hardening at 750-850° produces a more coarsely crystalline structure, and the metal becomes brittle with a tendency to develop hair-cracks. The hypereutectoid steel requires very slow cooling after annealing at 900° and a hardening temp. of 800-900° to produce a tough structure. Lamellar cementite causes brittleness and hair-cracks to develop on hardening, whereas granular cementite or troostite imparts a high degree of toughness to steel.

Temper hardening of steel containing copper. H. EICHENHOLT AND W. KESTER. *Steel & Iron* 50, 687-90 (1930).—The solid soln. of Cu in α -Fe is 0.4% at 0-600° and then increases according to the equation, $\log x = -4125/T + 4.22$, to 3.4% at the

eutectoid temp., 810° . On cooling a Cn steel from the γ -solid soln range, supersatd. solid solns of Cu in α -Fe, having a structure resembling that of martensite, may be obtained; these supersatd solns., as well as those contg 0.0-3.4% Cu quenched at 600 - 800° , undergo age hardening on tempering at 450 - 600° , but the sepn of the excess Cu in a highly dispersed condition and the subsequent coagulation of the particles take place very slowly. When the steel contains both Cu and C in supersatd solid solns., both elements sep in a disperse form on aging at high temps., the 2 reactions taking place independently of one another. Thus on tempering a quenched steel with 0.05% C and 2% Cu at gradually increasing temps., the hardness rises sharply from 160 to 210 at 20° during 3 days, then to 220 at 100° because of sepn of excess C, the hardness then falls rapidly to 150 at 300° , because of coagulation of the cementite, and rises again to 205 at 500 - 550° because of pptn of the Cu. The properties of temper-hardened forgings of steel with 0.8% Cu and 0.4% Cr resemble those of a Cr-Ni steel with a tensile strength of 70 kg/sq mm., both steels are characterized by the regularity of their mech. properties throughout large cross sections and by their stability on prolonged storage after tempering. B. C. A.

The malleability of rolled soft steel and a new method for its improvement. JOSEPH VITÓRISZ. *Iron and Steel Inst. (London) Carnegie Scholarship Mem.* 19, 167-213 (1930).—Tensile tests and microscopical investigations were performed on a com. open hearth mild steel contg C 0.05, Mn 0.29, P 0.02 and S 0.035%. Specimens obtained from the same billet (heated to 1350° and rolled in 8 passes at 1200°) were rolled at temps. ranging from 1250° down to 20° . Test pieces were then taken both in the direction of rolling and across. The following quantities were detd. for various degrees of deformation: yield point, traction hardness, uniform and total elongation, tenacity, stress in contracted section of fracture, malleability, reduction in area, tensile strength and no. of grains in a given field of vision. The changes in properties were recorded in graphs as a function of the rolling temp. The main results are summarized as follows. Between 700° and 850° all mech. properties have very low values, this range may be taken as the dividing line between hot- and cold rolling. Uniform and total elongation are at their max. value on rolling at 900° and reach a min. at 300° . Rolling between 200° and 400° gives a fairly const. max. value for yield point, traction hardness and tensile strength. At const. temp. these quantities, as a rule, increase when the degree of deformation increases. The tenacity curve shows a rise at 850° , falls rapidly between 850° and 750° , then rises again at 700° and drops to its lowest value at 300° . The grain no. increases rapidly with decrease in rolling temp., to 850° , but below 700° no appreciable change in grain no. is found. After rolling between 200° and 550° a special etching method with the aid of Fry's reagent brings out characteristic slip figures at angles of 45° to the direction of compression parallel to the direction of rolling and in the form of horizontal lines across the direction of rolling. Rolling at room temp. produces merely slipbands, serrated crystal boundaries and broken crystals. It is suggested that recrystn. of specimens rolled between 200° and 400° will produce the finest structures and the best mech. properties. This tentative conclusion is contrary to the customary practice of recrystallizing material rolled at room temp.

H. S. V. KLOOSTER

Carburization and decarburization of carbon steels. TAKEJIRO MURAKAMI AND HARUJIRO SEKIGUCHI. *Tetsu-to-Hagane (J. Iron and Steel Inst. Japan)* 16, 1015-40 (1930).—C steels with 0.29, 0.70, 1.02 and 1.29% C were treated at 800° , 900° and 1000° with different kinds of carburizing agents. With pure C powder, decarburization was already noticed at 800° in steels whose C content was higher than 0.3, especially when the packing was insufficient. Carbonate salts prevent this effect, though slight decarburization was found at initial stages. Carburizing speeds can be accelerated by the addn. of BaCO_3 till the amt. of the salt reaches 40%; the addn. of NaCO_3 to this mixt. promotes the tendency. When the carburizing temp. exceeds 900° fine particles of C adhere to the steel surface. It is easy to carburize up to the Ac₁ point, but the carburizing speed drops suddenly above this. Small quantities of moisture in C are preferable for carburization, and the size of C powder has no effect on the speed.

M. KURODA

The effect of carbon on the transformation points and hardness of 12% chromium steels. TAKEJIRO MURAKAMI AND YOSHIRO FUJII. *Kinzoku-No-Kenkyu (J. for Study of Metals)* 7, 505-23 (1930).—By measurements of the thermal dilatation and the magnetization of 8 specimens contg. about 12% Cr with const. and varying C contents of less than 2.04% during heating and cooling under different conditions, the change of transformation points of these specimens was studied. The effect of C on the A₁, A₂ and A₃ points in slow heating and cooling from 900° and 1100° with different ve-

locities was carefully observed, and the results were shown in curves. The Brinell hardness of specimens cooled under the same varying conditions as above, or quenched in oil from 900°, 1000°, 1100°, and 1200°, or quenched and tempered at several temps. was measured. From the results, the change of hardness with the C content and with the several heat treatments was plotted in several curves and discussed in detail. The microstructure of these specimens and its change under several heat treatments were also described. M. KURODA

Manganese steel. HEINRICH KARUTZ VON SCHREIBL *Giesserei-Zig* 27, 93-100 (1930)—A review of the prepn of high Mn steel castings. CURTIS L. WILSON

List of alloys. WM CAMPBELL. *Proc Am Soc Testing Materials* 30, Pt 1, 336-97 (1930)—Campbell's list of non ferrous alloys, which has been so useful for reference, has been brought up to date, and greatly enlarged. There are about 8000 compns listed. The phys properties of the standard and widely used alloys are given from the Am Soc for Testing Materials specifications and other sources. H. W. GILBERT

Comparison tests on light metal piston alloys. M. V. SCHWARTZ. *Z. Metallkunde* 22, 417-9 (1930)—An app is described for detg the performance of a piston in service. Data obtained with it are used to compare Al alloys and cast Fe used in piston manif. ROBERT F. MEIN

Measurements on the superconductivity of alloys. J. C. McLENNAN, J. F. ALLEN AND J. O. WILHELM. *Trans Roy Soc Can.* [3], 24, Sect 3, 53-64 (1930), cf. McLENNAN and QUINN, *C A* 25, 2050.—The supercond transition points of a no of alloys contg Sb, As and Sn were measured to compare the effect produced with that of metals from other groups. Most metals when alloyed with superconductors do not affect the supercond appreciably. Au apparently lowers the supercond transition point of some metals while with one exception the metals of the 11 group elevate the transition temp of the metal with which they are alloyed. Tables giving the transition temp for the alloys examd are included. J. W. SIMLEY

Methods of mechanical testing of light aluminum alloys used in automobile and aircraft construction. S. S. NEWART. *1st Communications New Intern Assoc for Testing of Materials A*, 277-92 (1930)—Investigations were carried out to det. the methods of testing which best show the influence of the following factors: heat treatment, the addition of other metals (Mg, Fe) to the alloy, and methods of casting. Tensile strength, compression strength, elongation and the results of a bending test with a single blow are not affected by these factors. On the other hand, the following tests are sufficiently sensitive to enable the mech properties to be estd and to det the influence of the factors under investigation: bending strength, limit of proportionality under compression, limit of endurance (fatigue test), rupture under repeated blows, alternating tensile and compressive stresses by repeated blows and shortening under compression. A definite relation exists between the results of these tests and the elements of construction. The best method for detg the resistance to corrosion is the soly test for 15 min in a 10% soln of HCl. There is a relation between the results of this test and those of the fatigue tests. No difference is observable between the mech properties of specimens cut from the main casting and those of specimens cast separately. Al alloys can be hardened even if they contain no Mg. The addn of Mg up to 0.6% and of Fe up to 1.0% increases considerably the hardness and the strength after hardening. Like all other metals, Al alloys have a safe limit of fatigue stress. E. J. C.

The notch toughness of a few aluminum alloys especially at low temperatures. W. A. GOLDNER. *Z. Metallkunde* 22, 412-6 (1930)—Conclusion of a previously published article (cf *C A* 24, 5009). The influence of velocity of test, of the width of the test piece, and of the depth of the notch, all at -180° were studied. The relations between these conditions is expressed mathematically and graphically, and discussed at length with comparison to Fe alloys. ROBERT F. MEIN

Aluminum alloys containing nickel. W. C. DAVERAUX. *Brass World* 27, 31-6 (1931)—The manuf and applications of "Y" and "RR" alloys are outlined. "Y" metal contains 4% Cu, 2% Ni, 1.5% Mg, balance Al; "RR" contains 1.3-2.25% Cu, 1.3% Ni, 0.1-1.6% Mg, 0.1-0.8% Ti, 0.5-2.2% Sn, balance Al. The manuf of Cu hardener, Ni hardener and Cu-Ni hardener is also outlined. Applications of "Y" and "RR" alloys include aero and automobile engineering, cylinder heads and pistons, connecting rods, propellers, crank cases, bus frames, etc. W. H. BOYNTON

The advantages of titanium additions to aluminum. H. RÖHRIG. *Metallwirtschaft* 10, No 6, 105-11 (1931)—It is difficult to alloy Ti with Al on account of the high m p of Ti. The best method is to add TiO₂ to the Al ore and reduce the 2 together electrolytically. Ti added to Al in small quantities forms either TiAl₃ or TiAl, a

brittle compd. Pure Al has a very coarse grained structure. By alloying with small quantities of Ti, 0.01 to 0.03%, the grain size is reduced considerably. This is illustrated by micrographs. Sheet Al, 1 mm thick, contg 0.01-0.19% Ti, was annealed at 500°, stretched cold 2, 5 and 10% and then annealed at 600°. The samples were etched in HCl-HF to show the recrystn structure. With the higher Ti contents the grain size was smaller, but the difference was less noticeable with the 10% cold-worked samples. With Ti free Al recrystn begins at 300°, with 0.19% Ti at 350°. The influence of these small quantities of Ti on the tensile strength and ductility is slight. The tensile strength of the samples contg Ti which were hard rolled, and those which were annealed at 350°, was slightly higher and the ductility lower. For those annealed at 450° and 500° the reverse was the case. Corrosion tests were made by immersing samples for 1 hr at 20° in 4 parts HNO₃, 4 parts H₂O and 1 part HCl. The loss in weight increased in the following order: annealed at 500° (least), annealed at 450°, at 400°, at 350°, hard rolled (most). Samples contg Ti lost less weight than those without Ti. Those which were annealed during the hard rolling process lost more weight than those not annealed. C E MACFARLANE

Study of iron-aluminum-carbon alloys. O V KEIL AND O JUNGWIRTH. *Arch Eisenhüttenw.* 4, 221-4 (1930). —The most suitable melting methods for the production of Fe-Al-C alloys are detd. Up to 14% Al, the eutectic temp rises from 1145° to 1300° and then slowly falls again. With 1% Al, the eutectic point is reduced by about 0.16% C. The greatest formation of graphite lies between 3 and 20% Al, between 12 and 18% Al the melt when chilled is white. With increasing Al content 2 new constituents are formed: Phase I, which is low in C, microscopically homogeneous but easily sol in acid; and Phase II, probably an Al-Fe carbide with great stability against acid attack. H STOERTZ

Electrical conductivity of some bismuth alloys at low temperatures. J C. McLENNAN, J F. ALLEN AND J O WILHELM. *Trans Roy Soc Can.* [3] 24, Sect 3, 25-35 (1930). —A no. of alloys contg Bi were prepd and their super-conducting transition temps detd with an accuracy of about 0.1°K. The transition temps found were: 58% Bi and 42% Pb 8.8°K., Rose's metal (Bi₃SnPb) 8.5°K., Newton's metal (Bi 50, Sn 19, Pb 31%) 8.5°K., Wood's metal (Sn 12.5, Cd 12.5, Pb 25, Bi 50%) 8.2°K., Bi₄Tl₃ 6.4°K., Pb 83 and 17% Sb 6.6°K. The presence of Bi in an alloy tends to raise the superconducting temp while the presence of Sb sometimes raises and sometimes lowers the superconducting point. The Pb-Bi alloy has a superconducting point 1.6°K. higher than any other previously obtained. Attention is drawn to the inadequacy of all the theories of elec. cond in metals in that they do not account for the phenomenon of superconductivity. A suggestion is made that superconductivity may be due in part, at least, to the directive action of magnetic forces within the atom. J W SHIPLEY

Constitution of the cadmium-rich alloys of the system cadmium-silver. P. J. DURRANT. *J. Inst Metals*, Advance Copy, No 549, 15 pp (1931). —Alloys contg up to 40% Ag were studied microscopically and by means of heating and cooling curves. There are 3 peritectic lines, and 3 solid solns in this area. The solid soly of Ag in Cd is 6.0% at 343°, and 5.2% at 250°. The second solid soln includes 18-33.5% Ag, the third more than 36% Ag. Also in *Metal Ind* (London) 38, 403-6, 415 (1931). J L GREGG

Copper alloy systems with an α -phase having variable limits and their use in the hardening of copper. M. G CORSON. *Rev met* 27, 83-101, 133-53, 194-213, 265-81 (1930). —An account is given of the investigations leading up to the development of the so-called "Corson alloys" which have been the subject of numerous patents in recent years. The alloys consist of Cu with small percentages of Cr, Co, Si, or silicides of Ni, Co, Fe or Cr, all of which constituents have a higher solid soly at high temps. than at the ordinary temp. Between 20° and 1000° the solid soly of Cr in Cu varies from 0.05 to 0.8%, of Co from 0.35 to 3.4%, of Co₂Si from 0.3 to 3.3%, of Ni₂Si from 0.7 to 8.2%, of Fe from 0.2 to 3.8% and of Si from 2.8 to 7.5%. Alloys contg silicides are best hardened by quenching from 800° to 1000° and aging at 250-450°, these alloys should contain a slight excess of the 2nd metal over that required to form its silicide so as to insure max. hardening on aging without the serious fall in elec. cond, which is produced by the retention of Si in solid soln. Photomicrographs of numerous alloys of this type are reproduced, and the effect of heat treatment on the mech and elec. properties and on the structure is illustrated by graphs and photographs. B C A

X-ray investigation of copper-arsenic alloys. NOBUYUKI KATOH. *Z. Krist.* 76, 228-34 (1930) (in English). —See C. A. 25, 272. L S RAMSDELL

Investigation of the effects of impurities on copper. VII. The effect of antimony

on copper VIII. The combined effect of antimony and arsenic on copper. S. L. ARCHBUTT and W. E. PRYTHIECH *J Inst Metals*, Advance copy, No 546, 33 pp (1931); of *C A* 24, 2705.—Alloys contg less than 0.02% O were investigated. Those contg up to 0.47% Sb rolled hot, while one with 0.85% Sb was hot short. Annealed ingots contg up to 0.85% Sb could be cold rolled. All alloys contg 0.47% or less Sb could be rolled into thin sheet. Sb increases the strength without appreciably affecting the ductility. The endurance limit of the annealed 0.47 Sb alloy was 7.8 tons/sq in as compared with 5 or 6 tons/sq in for pure Cu. The softening temp is raised by Sb, its effect being greater than that of As. The limit of soly of Sb in solid Cu was detd microscopically as approx 10%. The addn of As to alloys contg Sb had only a slight effect on the properties. All of the alloys had high impact strengths. J. L. G.

The equilibrium diagram of copper-tin alloys. MATSUJIRI HAMASUMI and SEIJI NISHIGORI *Kinzoku No Kenkyu (J for Study of Metals)* 7, 535-51 (1930).—There are many structures of Cu-Sn alloys (especially those contg 10-40% Sn) that are hardly explainable by any of the existing diagrams. The latest diagram is probably that of Raper. H and N prepd bar specimens of 15 cm length and 5 mm diam in spite of their glass-like brittleness and measured the change of elec resistance and the thermal dilatation on heating. Thermal analyses and microscopical exams on the quenched specimens were made, and the equil of the system in the solid state was detd. There are 2 intermetallic compds, namely, δ and η in this region, corresponding to the chem formulas Cu_3Sn , Sn_2 , and Cu_2Sn . The compd δ formerly considered to have the formula Cu_3Sn , is proved to be Cu_2Sn , as pointed out by Westgren. The compound η undergoes a transformation at 675° to the solid soln γ . This transformation accompanies an expansion of 0.56% of its linear dimension, which is twice as great as that of γ iron to α iron. There are 3 eutectoid transformations at 510°, 570° and 630°, corresponding to the equil $\alpha + \delta \rightleftharpoons \beta$, $\delta + \eta \rightleftharpoons \epsilon$ and $\beta + \eta \rightleftharpoons \gamma$, and 2 peritectic transformations at 580° and 625°, corresponding to the equil $\delta \rightleftharpoons \beta + \epsilon$ and $\epsilon \rightleftharpoons \beta + \eta$. The transformation of δ and η into ϵ accompanies a linear expansion of 0.25%, which is nearly equal to the expansion of γ iron into α -iron. The typical dilatation temp and resistance temp curves of the samples are shown. The change of the direction of the curves on heating exactly coincides with the equil diagram given. The existence of the new phase ϵ is also clearly proved by the microscopic exam of the quenched specimens. M. KURODA.

X-ray examination of the lattice structure of the ϵ phase of the copper-tin system. J. O. LYONS *Ann Physik* [5], 8, 124-8 (1931).—An alloy with exactly 25 at % Sn was prepd from pure materials, melted in a H atm and carefully homogenized. After powdering in an agate mortar it was heated in an evacuated glass tube for 40 hrs at 400° and then subjected to the x-ray exam with an angle range of 25-75°. Extra lines, not formerly observed, were found, which could not be identified as lines of the hexagonal lattice with only 2 atoms in the elementary range. These lines were found to be surprisingly strong by choosing an 0.2 mm camera opening and an exposure of 17 hrs. Being dissolved into their α doubles, they could be measured very well. The indices of all the lines observed could be detd. These extra lines found indicate a regular distribution of the 2 kinds of atoms in the lattice, which is in agreement with the low elec resistance of Cu_2Sn . It was not possible to find a simple at. arrangement, with hexagonal symmetry, which would explain all the lines observed. The lattice appears to be very complicated. LEOPOLD PESSEL.

Alloys of iron, nickel and chromium. M. A. GROSSMAN *1st Communications New Intern Assoc for Testing of Materials A*, 97-103 (1930).—A review of the properties and uses of these alloys. Fe-Ni alloys are used with various proportions of Ni to give high elec resistance, low thermal expansion, high permeability and to form permalloy (78% Ni). Fe-Cr alloys are used for their stainless properties and stiffness at high temps. Ni-Cr alloys (80-90% Ni, 20-10% Cr) give high elec resistance combined with resistance to oxidation and are therefore used for elec. heating elements. Fe-Cr-Ni alloys are rustless, readily worked and have a high creep stress. Alloy steels may contain from 0.5 to 4% Cr, giving great hardness, or from 0.5 to 5% Ni to give toughness. Ni and Cr may also be combined with other alloying elements. E. J. C.

Tensile properties of alloy steels at elevated temperatures as determined by the "short-time" method. WILLIAM KARLBAUM, R. L. DOWDELL and W. A. TUCKER *Bur Standards J Research* 6, 199-219 (1931).—The materials tested were a plain carbon steel and commercial alloys of Cr, V and Fe with and without addns of W, Si or Al, also Cr-W, Ni-Mo and several austenitic steels with and without W. Metallographic study revealed very little intercryst. weakness and no marked or significant change in structure resulting from the high temp tests. The addn of W increases the "structural

stability" of the austenitic alloys. Hardness was detd as a further indication of tempering which might have occurred during the heating. A comparison of short-time test and "flow test" for pearlitic and austenitic steel is included. R. R.

Mechanical and physical properties of magnesium alloys. S. L. ARCHBUTT. *1st Communications New Intern Assoc for Testing of Materials A*, 251-6 (1930)—A review of the results of the investigations carried out during recent years on Mg and its alloys. Important progress has been made in regard to purity of the metal and foundry technique, but intensive research is necessary regarding heat treatment and modification processes. Al, Zn and Cu are the chief alloying elements at present in com use, and curves are given showing the tensile strength and ductility of binary alloys of Mg with these metals as well as with Ni. Four tables are also included showing the results of numerous tests carried out by various bodies on Mg alloys. E. J. C.

Compression diagrams and temperature-hardness curves of some lead-tin-antimony alloys. EDUARD HERTEL AND ARTHUR DEMMER. *Metallwirtschaft* 10, No. 7, 125-6 (1931)—Compression stress-strain diagrams were prepd for 12 Pb-Sn-Sb alloys, with samples 20 mm diam and 25 mm high the load was applied at the rate of 0.01 mm per sec. These results and Brinell hardness tests are shown in the tabulation. In 10 x ray exams, the existence of a tetragonal Sn-rich, a trigonal Sb-rich and a cubic constituent was found and the absence of any definite Sn-Sb compds was proved.

No	Sn	Pb	Compu Sb	Cu	Ni	Mg	P	0.2% Load	0.4% Load	0.9% Load	1.8% Load	3.6% Load	Brinell hardness at 20°	Brinell hardness at 100°	Ratio 0.4% 0.2%
	100							2 51	2 94	3 40	3 7	4 03	12 8	5	1 17
		100						2 07	2 58	3 08	3 39	3 69	6 1	3 8	1 24
			100					5 31	8 64	12 54	Broke		49 0	34 5	1 06
				100									63 0	57 0	
I	50	50						2 57	4 09	4 77	5 20	5 49	14 5	6 8	1 59
II	48	48	4					3 42	4 59	5 56	5 81	5 86	18 6	8 0	1 34
III	44	44	12					4 26		7 00	7 47	7 77	24 9	9 5	
IV	42	42	16					4 36	6 04	7 93	8 51	8 84	28 4	13 0	1 36
V	40	40	20					4 87	6 53	8 86	9 75	10 19	30 0	13 5	1 33
VI	40	40	15	5				5 11	7 33	9 89	10 45	10 87	43 0	22 8	1 43
VII	40	45	10	5				4 44	6 87	9 39	9 98	10 23	30 4	12 4	1 54
VIII	40	40	15	3	2			4 78	7 10	9 89	10 40	10 67	29 7	15 9	1 48
IX	40	40	12	3	5			2 53	3 10	5 17	6 71	6 87	31 2	17 2	1 21
X	40	40	9	3	8								34 5	17 8	
XI	40	40	15	3		2		4 97	6 20	8 12	8 86	0 30	1 24
XII	42	42	10	2		2 2		4 40	6 16	7 21	8 01	8 62	1 40

C. E. MACFARLANE

Silver-cadmium alloys. E. R. TREWS. *Deut. Goldschmiede-Ztg* 33, 56-8 (1930).

B. C. A.

Alloys of zinc and manganese. N. PARRAVANO AND V. MONTORO. *Mem. accad Italia, Classe sci fis. mat e nat* 1, *Chim.* No. 4, 19 pp.—A study of the binary alloy Zn-Mn by the x ray method. Alloys studied contained up to 23.3% Mn. In these alloys 3 different phases were detd: (1) a solid soln of Mn in Zn, the α phase, with hexagonal lattice that is similar to that of Zn, having up to 0.93% Mn, (2) an ϵ phase with hexagonal lattice different from that of Zn. This corresponds to the compd. $MnZn$, which contains 16 atoms ($2MnZn$), (3) a γ phase, with body-centered cubic lattice, corresponding to the compd. $MnZn_2$ and having 48 atoms ($12MnZn_2$) per unit lattice. G. T. MOROX

The influence of third metals upon the constitution of the brasses. III. The effect of tin. O. BAUER AND M. HANSEN. *Z. Metallkunde* 22, 405-11 (1930).—Continuation of previously published articles (cf C. A. 25, 1202). The exptl. methods are described and the results expressed in diagrams representing sections of the ternary system. The results are summarized in a partial ternary three-dimensional diagram.

ROBERT F. MEHL

Influence of the degree of rolling on the properties of α -brass. N. DAVIDENKOV AND V. BUGAKOV. *Metallwirtschaft* 10, 1-6 (1931).— α -Brass (62.95% Cu, 37.05% Zn) was rolled to thicknesses corresponding to reductions varying from 0 to 80%. The tensile strength, elastic limit, elongation, Brinell hardness, the soly in HNO_3 , and the micro- and x ray structure in the rolled condition were investigated. Between 20 and 30% reduction there is a crit. region where all curves of properties plotted against % reduction show discontinuities. On annealing at 200° the change in proper-

ties becomes continuous, showing that the discontinuities are caused by internal stresses C. H. LORIO

Deep-etch test of brass. W. F. GRAHAM AND L. A. MARISSSE *Trans & Bull Am Foundrymen's Assoc.* 2, 810-25(1931)—Deep etching of brass with a soln of HNO_3 (1.42 sp. gr.) and 10% by vol. of HCl (1.18 sp. gr.) clearly brings out areas of discontinuity of the structure caused by solms, crystal boundaries or dendrites C. H. L.

Some properties of silicon "Al-bronzes." L. J. HAYEN *J. Inst. Metals, Advance copy*, No. 547, 18 pp (1931)—The effect of Si on the mech. properties of Cu base alloys contg. 5, 7½, and 10% Al was detd. The alloys were tested as cast, as cast and annealed by heating to 800° followed by air cooling, as quenched in oil from 800° or 600° and a few as forged. Si contents were as high as 4.5%. With 5% Al the addn. of up to 2% Si improves the properties. Further increase of Si increases the strength but lowers the ductility rapidly. An annealed 5% Al, 3.5% Si alloy has the properties of a 10% Al alloy. The addn. of up to 4% Si to the 7½% Si alloy progressively increases the strength and decreases ductility. With more than 3% Si the ductility is so low as to render the alloy useless. The addn. of even 1% Si to the 10% Al alloy produces a very low ductility. A microscopic study of the alloys is included. Also in *Engineering* 131, 498-500(1931) J. L. GREGG

Repairing pewter. A. EYLES *Brass World* 27, 27-8(1931)—The handling of pewter in the old days is outlined and practical notes are given for the reconditioning of old pewter vessels. In repairing old pewter vessels, dents must be removed and all surfaces must be perfectly clean. A good solder consists of Bi 50, Sn 25 and Pb 25%. Fluxes are rosin, Gallipoli oil, olive oil and a mixt. of rosin and olive oil W. H. B.

A sand-cast test bar for specification purposes for aluminum alloy castings. S. L. ARCHBUTT *1st Communications New Intern. Assoc. for Testing of Materials A*, 273-6 (1930)—Investigation into the use of a sand-cast bar prepd. in sand molds of special and simple type, with V-alloy of widely varying gas content, is described. Density, machining and tensile tests on bars so prepd. are compared with corresponding values obtained from chill-cast bars of similar dimensions. The presence or absence of gas is clearly indicated in the results from the sand-cast bars, while relatively little effect is seen in the tests on chill-cast bars L. J. C.

Magnesium alloy castings. C. FLAYEN *Metal Ind (London)* 38, 7-10, 31-2 (1931)—A review is given of the properties and foundry practice. The application of a special flux Elrasal which is to prevent the contamination of the metal is described, as well as the use of powd. Si in the core material, to obtain clean castings. L. P.

The mechanical work of cutting and dressing rough castings. M. KURZEN *Z. Metallkunde* 22, 382-6(1930)—A discussion is given of the evaluation of the mech. work performed, and its relation to design, by chisels, milling cutters, saws, grinding wheels and sand blast. The article is too detailed to permit adequate abstracting ROBERT F. MEHL

Some notes on a "burned-on" casting. L. W. BOLTON *Bull. Brit. Cast. Iron Research Assoc.* 1931, No. 31, 342-3—Bureau Report No. 68 describes physical and microscopic tests of a gray iron casting 1.5 in. thick, which had an almost circular hole about 3 in. in diam. filled in by burning on new metal R. RIMBACH

Some new facts of a chemical nature in the field of corrosion investigations and in the protection of metallurgical products against corrosion. JEAN COURCELOT. *Bull. soc. chim.* 47, 802-25(1930)—A review is given. Among the protective processes mentioned, the electroplating of Ni, Cd, Cr and parkerizing are discussed in greater detail LEFOLD FESSEL

Two years' progress in corrosion resistance. JOHN A. MATHEWS. *Paper Trade J.* 92, No. 10, 57-8(1931)—A brief review is given of some of the causes of failures of 18-8 alloys, particularly in the sulfite industry. The latest metallurgical theories are given relative to C ppts., and the need for proper heat treatment of metals is stressed, as well as the chem. compn. of the alloys. Means of preventing metal embrittlement is indicated A. PAPINEAU COUTURE

The resistance of electrodeposits to corrosion, with special reference to cadmium and zinc. W. S. PATTERSON *Metal Ind (London)* 36, 527-9, 579-82, 632(1930)—Of the 3 different types of corrosion of electrodeposits only the 1st is discussed—the corrosion of the coating metal considered as a continuous envelope sealing completely the metal it is protecting. The effects of impurities and of porosity of deposits upon corrosion were studied under a variety of conditions. Electrodeposited and hot-galvanized Zn are compared. Cd is regarded as an excellent coating for indoor work or for mild conditions of exposure, but for outdoor work where rain and atm. pollution

play an important part Cd fails and Zn is the more satisfactory and durable

W H BOYNTON

Corrosion resistance of light alloys. H. SUTTON. *1st Communications New Intern Assoc for Testing of Materials A*, 304-8(1930)—A review of the more important work of British investigators. E J C

The corrosion of aluminum casting alloys. R. STEAVER RAINER. *Z Metallkunde* 22, 357-61(1930)—The sensitivity of Al casting alloys toward corrosion is dependent upon the compn and upon the corrosive agent. Contrary to the usual opinion that the corrosion of Al alloys cannot be improved by addns, it has been shown that all those substances which are effective in sepg Fe and Si are effective in improving the corrosion resistance. It is not certain that a protective coating forms on Al alloys inhibiting further action, though the effect of small addns of Sb and Bi furnish some support for such a belief. Other addns decrease the corrosion resistance, especially Cu. The application of special alloys to special conditions is discussed. R F M

Corrosion resistance of aluminum and its alloys. M BOSSHARD. *1st Communications New Intern Assoc for Testing of Materials A*, 293-303(1930)—The great resistance which Al offers to the action of air, water and aq salt solns is due to the formation of a skin of oxide. Corrosion must be divided into 2 classes according to the way in which it occurs. (1) general corrosion, soly, (2) local corrosion, formation of holes. General corrosion occurs only under the action of corrosive agents which dissolve the oxide. On the other hand, corrosive agents causing local corrosion have at first no oxide dissolving qualities. The ability to dissolve the inert surface layer locally is a secondary effect of electrochem action. The effect of impurities (Fe, Si, Cu, Mg, Zn and Mn) on the corrosion of Al is also discussed. E J C

Studies of corrosion in aluminum and aluminum alloys under the microscope. H RÖNNRÖ. *Z Metallkunde* 22, 362-4(1930), cf *C A* 24, 1835—From microscopic studies reported R. concludes (1) that deliberate or accidental inclusions are more noble to corrosive attack than the Al matrix, (2) that pores or voids offer assistance to the progress of corrosion, (3) that according to the nature of the attacking substance corrosion may proceed intergranularly or along strings of inclusions, or it may proceed concentrically from the starting point without regard for structure—the first is typified by the attack of sea water or sea water accelerated by H_2O_2 , the second by boiling phenol or dil H_2SO_4 . The purer the Al the fewer are the inclusions from which attack might start. Corrosion resistance may also be improved by heat treatment.

ROBERT F MEHL

Corrosion of early Chinese bronzes. WM F COLLINS. *J Inst Metals*, Advance copy, Nn 348, 25 pp (1931)—The probable mechanism of the formation of patina on early Chinese bronze is discussed in light of the corrosion studies made in England and the mechanism of mineral formation. S was probably not a factor in the corrosion of the Chinese bronzes. The ancient Chinese bronzes are unique in contg appreciable quantities of Pb. The analyses of a large no of these bronzes are given. The data also include the probable period of origin. J L GARRIC

Corrosion of steel water pipes by stray electric currents. OTTO ROTHIE. *Rev brasl chim* 2, 249-54(1930)—After a period of 1-2 years, pipes showed holes which from appearances and the presence of FeCl_2 at places not yet perforated made the deduction reasonable that the damage was caused by continuous elec current. R tested pipes at various points of the city and found that they were positively charged in relation to the soil around them, the p d from the voltage of the trolley car power lines had a max. of 10 v. It was possible to reproduce in the lab funnel-shaped holes full of FeCl_2 similar to the ones in the city water pipes. The perforations occur where the asphalt protection is damaged and the elec current is permitted to pass. A diagram is given. JOHN M LADINO

Experiments with coal-fired pot-annealing furnaces. H STÄBLER. *Stahl u Eisen* 50, 381-91(1930)—An attempt has been made to elucidate some questions regarding annealing time, fuel consumption, choice of pot, etc., arising from an analysis of the details of operation of annealing furnaces in a no of works. The max and min. temps. within the heating stock, in a semi gas-fired furnace capable of accommodating 5 pots, have been detd as a function of the time. The annealing time, s. e., the time required for the min. temp within the material to reach the arbitrarily chosen value of 700° , is a linear function of the wt. of the charge (W) and is given by the equation $T = T_0 + KW$, where T_0 is the time required for the inner wall of the empty pot to reach 700° . The temp difference, $t_{\text{max}} - t_{\text{min}}$, when $t_{\text{min}} = 700^\circ$, increases linearly with the wt of the charge until the latter reaches about 1500 kg, when it passes through a max. and thereafter decreases. By suitably regulating the heating conditions, the char-

acteristic temp curves (t_{max} and t_{min} as functions of the time) can be varied to correspond with any desired heat treatment of the material. The calcn of the efficiency of the plant from the annealing time is discussed and illustrated by 1 or 2 examples. B. C. A.

Press welding of iron. HANS ESSEY *Arch Eisenhüttenw.* 4, 199-206 (1930) —E. detcs. the relationship between the press weldability of pure Fe or Fe-C alloys and temp welding pressure, surface properties and grain size. Three dimensional diagrams are given, in which tensile strength is plotted against temp and C content of press-welded Fe-C alloys. With pure Fe, the strength of the welded surface increases with temp. The smaller the grain size the more suitable is the material for welding, and at higher temps, the welding ability increases markedly with welding pressure. With increasing C content in pure Fe-C alloys, the temp range suitable for press welding becomes lower. H. STOERTE

The chemistry and metallurgy of Be (Drsow) 6. Sidelights on Cu as applied to brewing (STEWARTSON) 16. The melting point in the W-Re system (BECKER, MOERS) 2. Superstructure and magnetic susceptibility in the system Cu-Au (SEEMANN, VOOR) 2. Au and Ag solutions (Brit. pat. 338,383) 18. Clarifying suspensions [of ores] (Ger. pat. 517,055) 13. Apparatus for estimating the humidity of foundry sand (Ger. pat. 517,212) 1. Crystals of metals (U. S. pat. 1,793,672) 18.

FRENCH, HERBERT J. The Quenching of Steels. Cleveland: Am. Soc. for Steel Treating 177 pp.

MÖNZER, GEORG. Das Platin: Gewinnung, Handel, Verwendung. Leipzig: W. Diebener G m. b H. 136 pp. M 675

SCHER, JEAN: L'aluméc. Paris: Revue de l'aluminium et de ses applications. 11 pp.

Recovering ore values on an adhesive surface. ROYER LUCKENBACH. U. S. 1,792,544, Feb. 17. Candle tar with an oil thinner such as crude petroleum is used in sepp. ore constituents such as Au, Ag, Pt or other metals. Na silicate may be added to assist in sepp. of gang.

Apparatus for amalgamation of metal constituents in ores. EDGAR W. MYERS. U. S. 1,794,000, Feb. 24. Structural features.

Reducing iron from iron oxide ores. WM. W. PERCY. U. S. 1,792,507, Feb. 17. Preheated ore is charged into a reducing chamber and subjected to the reducing action of gases comprising CO and H₂, and after the gases have passed from the reducing chamber a portion of them is withdrawn. The temp of the remaining gases is maintained, and they are caused to pass downward through a bed of incandescent C in a bottom-blown gas regenerator exterior to the reducing chamber and thence return into the bottom of the reduction zone at a temp at which complete reduction of the Fe oxide will take place. App. is described.

Working up lead ores and waste. METALLGESELLSCHAFT A.-G. Ger. 514,679, April 29, 1927. Pb ores and waste contg. Pb, such as flue dust, are worked up in a completely rotatable cylinder furnace in which metals such as Zn, Cd and As are volatilized as oxides. The Pb is obtained, without the addn. of C, by the thermal reduction of PbSO₄ and PbS with production of SO₂ and SO₃. App. is described.

Working up ores containing zinc, lead and copper. ALBERT F. MEYERHOFER. Ger. 514,537, Dec. 24, 1924. A wet process for working up these ores consists in sepp. the Zn and Pb as sol. fluosilicates, leaving the Cu. The Zn and Pb fluosilicates are sepp. by fractional crystn. and worked up to obtain the metal with regeneration of HF. The Cu remains mostly as CuS and is worked up in the usual way. The detailed steps of the sepp. of the ZnSiF₆ and PbSiF₆ are given.

Treating zinciferous ores. A. FOLLIER and N. SAINDERICHIN. Brit. 337,636, Oct. 19, 1929. Volatile metals in Zn bearing ores contg. S are extd. by blowing air in 2 stages onto a thin traveling layer of the ore, first onto ore alone for desulfurization and then onto the bot. roasted product together with reducing C. App. is described. Cf. C. A. 24, 4499.

Tin from ores, etc. SOC. D'ÉLECTRO-CHIMIE, D'ÉLECTRO-MÉTALLURGIE ET DES ATELIERS ÉLECTRIQUES D'UGINE. Brit. 338,149, June 22, 1929. See Fr. 692,640 (C. A. 25, 1481).

Separating and refining metals. HIRSCH, KUPFER- UND MESSINGWERKE, A.-G. (Cyrano Tama, inventor). Ger. 514,612, Aug. 31, 1928. Metals and metallic compds.

of differing in p are sepd by placing them near the perforated bend of a tube and heating to the various fusing temps. of the metals or compds. by elec. induction.

Metal granules from molten metal. HARTSTORF-MIRAL A. G. Ger. 514,623, Dec. 6, 1928. Details of app. for atomizing the molten metal are described.

Floating metal powders. JOHN R. CARE (to Richardson Co.). U. S. 1,704,553, March 3. A dry powder metal such as Fe is treated with a volatile oily liquid such as benzene, and after the liquid is evaporated the powder is deposited on an aq. soln. of a metal salt such as a Cu salt soln., the metal of which will be replaced by the metal of the powder, and the resulting powder is consolidated into a bulb on the surface of the soln. The product thus formed is suitable for integrating into foils by the action of an elec. current.

Treating impure lead containing copper and arsenic. JOHN B. SCHWETZMEIER (to Bunker Hill & Sullivan Mining & Concentrating Co.). U. S. 1,704,617, March 3. For recovering Cu and As in the form of dross from hot impure lead to purify the lead, the initial dross is skimmed off from a molten mass of the material and the material of the molten mass is caused to continuously circulate (as by the action of a propeller submerged in the bath, in a described app.) downwards from and then upwards to the surface of the bath. A charge of NH_4Cl is introduced into the circulating molten mass to sep. the Cu and As from the Pb so that the first two will rise by upward currents and gravitate to the surface. The dross is skimmed off to recover the Cu and As from it while the mass is quiescent, the treatment with NH_4Cl and circulation of the material is repeated, and the temp. of the mass is gradually reduced as the process proceeds to freeze the Cu and As and facilitate their isolation by reason of their relatively higher m. ps. as compared with Pb, so that the dross will gradually become drier as the process proceeds.

Separation of gases such as sulfur dioxide from smelter fumes by adsorption under pressure. RETOUSET L. HAZEN and WM. H. DEXTER (to American Smelting & Refining Co.). U. S. 1,701,377, March 3. An adsorbent material such as silica gel may be used and during "dewarbing" sufficient heat is supplied to the material to balance the latent heat of vaporization of the gas and to maintain the temp. of the material substantially constant, so that the process is carried out substantially isothermally.

Ladle for melting metals (suitable for use in dental work). E. BRYNER, Brit. 338,003, Jan. 14, 1929. Structural features of a double walled ladle the jacket of which may be filled with glycerol or other suitable liquid.

Molding sand mixture for use in casting iron, steel, brass, etc. ROY CROSS (to Silica Products Co.). U. S. 1,703,011, March 3. Inert material such as sand and clay is mixed with about 2.5-3% of bentonite treated to increase its gelling power (suitably by conjoint use of MgO , Ca aluminate, Ca silicate, portland cement or blast-furnace slag).

Binder for mold materials such as sand. MONROE S. CLAWSON. U. S. 1,704,558, March 3. Sawdust, resin and paraffin oil are used together.

Coating molds such as those for casting automobile engine pistons, etc. NELSON J. QUINN (to Metal Castings Holding Corp.). U. S. 1,704,530, March 3. Various mech. details are described for applying a liquid contg. refractory material such as MgO to matrix surfaces from a nozzle first placed close to the surface to build up a relatively thick coating rapidly and then held at a greater distance from the surface to produce a surface coating which has the desired degree of porosity.

Ingot mold. HOWARD R. WILLIAMS (to Vulcan Mold and Iron Co.). U. S. 1,703,003, March 3. Ingot molds are made of cupola cast iron which contains less than 0.05% S and has a high resistance to cracking through sudden expansion.

Hot top for ingot molds. WM. H. RAMSON (to Votter Mould & Iron Corp.). U. S. 1,702,808, Feb. 17.

Hot top for ingot molds. NIKLAS F. FOLER (to Insulated Top Co.). U. S. 1,701,840, March 3.

Composite ingots. W. ALBERTS and F. STIRY. Brit. 338,022, Aug. 28, 1929. In casting ingots suitable for rolling into "vignoles" rails, etc., the partition for sepd. the molten metals forming the ingot is made of practically pure iron and may be perforated or woven.

Ingots for producing non-porous articles such as rock drill pistons, etc. LOUIS G. FIRTH (to Firth-Sterling Steel Co.). U. S. 1,702,361, Feb. 17. Metal such as steel is cast into an ingot having a body portion and flanges extending from the latter, the flanges are longitudinally sepd. from the body portion, and products such as pistons, valve seats or valves are formed from the flanges.

- Apparatus for die-casting metal articles. **FRANK F MANTEI** (to Friedrich J. Haas)
U S 1,793,110, Feb 17. Structural features.
- Casting metals. **HARRY ALBERT SCHWARTZ**. Ger 514,875, Mar. 4, 1924. App
for casting metals with cooling by a circulating liquid is described.
- Casting slab ingots from deoxidized steel. **FRIEDRICH GATHMANN**. U. S. 1,793,314.
Feb 17. App and various details for controlling solidification are described.
- Casting blooms of aluminum. **LIMES P McCLURE** (to Aluminum Products Co.)
U S 1,792,545, Feb 17. Mech features.
- Casting aluminum alloy pistons, etc. **FRITZ FEITENBACH**. U. S. 1,792,580.
Feb 17. Al is melted and Si is added in an amt. about 18% or somewhat more of the
metal mixt., and to this mixt are added particles of hard material such as Ag, Cu, B,
Be, Zr, Mg, Ti, V, Mn, Mo or Cr to approx. 5% of the ultimate metal mixt. The molten
mass is poured into a vertical mold and chilled, so that the hard metal particles are
distributed over the surface of the casting.
- Aluminum castings. **THEODORE LAMBLE**. Swiss 143,081, May 11, 1929. In mak-
ing Al castings, NaCl, KCl, K_2SO_4 , wood-charcoal dust, $Na_2B_4O_7$, $10H_2O$ and cryolite
are added to the molten metal. These additives prevent oxidation of the Al, adsorption
of N, H or CO_2 and remove Fe, Si, etc., as scum.
- Casting copper alloys. **ARTHUR KIRCHHOFF**. Ger 513,738, April 16, 1924. Thick
castings of Cu alloys are produced by adding CuO to the molten metal just before
casting in the proportions of 1 to 2 parts CuO to 50 parts of metal.
- Casting "stainless steel" or similar alloys. **W. E. MARTIN** and **J. A. BERLYN**
Brit. 337,413, July 29, 1929. Structural details are described of sand molds which
may be combined with chills and wire reinforcing in their walls.
- Alloys of tungsten or molybdenum carbides. **FRIEDRICH KAUFF A-G** (Hermann
Voigtlaender and Otto Kaufels, inventors). Ger 514,728, May 3, 1928. See Brit.
310,876 (C A 24, 584).
- Ore-roasting muffled-retort furnace. **HOWARD S. BAILEY** and **RAY A. BENNETT**
(Bailey to Bennett). U S 1,792,476, Feb 17.
- Calcine-dust seal for ore-roasting furnaces. **BENJAMIN G. CALL** (to American
Smelting and Refining Co.) U S 1,793,274, Feb 17. Structural features.
- Blast-furnace tuyère. **E. POHL, A. WAGNER** and **GUTENHOFFNUNGSGES. OBER-
HAEUSEN A-G**. Brit. 338,130, May 10, 1929.
- Tuyère for blast furnaces. **HOLMES B. GRONIGER**. U S 1,793,549, Feb 24.
- Drying air for furnace blasts. **G. MALYOTH**. Brit. 338,244, Aug 28, 1929. See
Ger 515,216 (C A 25, 1490).
- Multi-trough hearth furnace for distilling metals such as zinc, mercury and cad-
mium. **E. M. S. INDUSTRIAL PROCESSES, LTD.**, **R. A. STOKES** and **B. G. L. ROBERTS**.
Brit. 337,471, Aug 3, 1929. Structural features.
- Metallurgical furnace suitable for annealing, etc. **JOSEPH HARRINGTON** (to Joseph
Harrington Co.) U S 1,792,638, Feb 17. Structural features.
- Furnace suitable for continuous heating of metal sheet bars, slabs, billets, etc.
FRANK J. WINDA (to Allegheny Steel Co.) U. S. 1,793,037, Feb 17.
- Furnace for heating metal billets and sheets. **A. SMALLWOOD** and **J. FALLON**.
Brit. 337,302, June 25, 1929. Structural features.
- Furnace for heat treatment of small articles in baskets or other containers. **FRANK
T. COPE** (to Electric Furnace Co.) U S 1,794,151, Feb 24.
- Heating salt baths for heat treatment of metals. **ARTHUR E. BELLIS**. U. S.
1,792,674, Feb 17. A low voltage a. c. is continuously passed directly through the
bath material between an electrode immersed in the bath and the metal wall of the pot
or container.
- Heat treatment and quenching procedure for differential tempering of car wheels.
JAMES C. DAVIS. U S 1,794,445, March 3.
- Heat treatment of loaded electrical conductors. **V. E. LEOG** (to Electrical Research
Products, Inc.) Brit. 338,169 Nov 24, 1928. In making Cu conductors loaded with
magnetic material, with heat treatment, the Cu is first heated to about 1000° in the
presence of a reducing agent such as charcoal (in a closed vessel in a resistance furnace)
and afterward the loading material is applied, followed by further heat treatment and
annealing. App is described.
- Apparatus for conveying materials such as metal bars or sheets through heat-
treating furnaces. **LEONARD LAASON** (to Republic Steel Corp.) U S 1,793,940.
Feb 24. Structural features.
- Apparatus for annealing arched metal sheets. **WM. E. WATKINS** (to Copper Plate
Sheet & Tube Co.) U S 1,794,659, March 3. Structural features.

Iron. WALTER G. CLARK. U. S. 1,792,967, Feb. 17. For producing a substantially C free metal, iron is melted and heated air is blown through it until about 75% of the Fe is converted into oxide, the residual air is displaced with N and heated H is then blown through the mass until the oxide is reduced to metal and until the C content of the metal is reduced to the desired limit.

Cast iron. EDWARD R. WILLIAMS (to Vulcan Mold and Iron Co.) U. S. 1,793,268, Feb. 17. For making cast iron with a pearlitic structure, a charge of low Si and C contents is melted, the S content is reduced to below 0.05%, the molten metal is poured into a mold which is at room temp. and the casting is cooled normally without application of external heat.

Cast iron. MEEHANITE METAL CORP. Brit. 337,844, Aug. 13, 1929. Iron which has been produced by methods such as are described in Brit. 210,118 (C. A. 18, 1640), Brit. 292,164 (C. A. 23, 1382) and Brit. 312,126 (C. A. 24, 819) is heat treated by maintaining the casting at a temp. above the A₁ point (about 800°) for over 12 hrs. and then cooling (the first stage of cooling—as from about 900° to about 540°—being at a rate not exceeding about 10° per hr.)

Iron and iron alloys. T. D. YENSEN and N. A. ZIEGLER (to Associated Electrical Industries, Ltd.) Brit. 338,409, Jan. 18, 1929. For obtaining iron or alloys of Fe with Ni or Si, substantially free from C and O, there is added to the material such a quantity of either C or O that these elements are present in substantially equal atomic proportions and the material is melted while simultaneously removing the products of combustion as they are formed (as by evacuation or passing an inert gas through the furnace). C may be added as such and O may be added as free O, air or an Fe oxide or oxide of an alloying ingredient. The product may be allowed to solidify in a vacuum and cooled slowly from 900–1200° to normal temp. (with previous hot working if desired).

Heat treatment of iron and steel sheets. HERBERT M. COX and CARY M. SMITH (to Republic Steel Corp.) U. S. 1,792,573, Feb. 17. The sheets are heated to a temp. near the lower crit. point of the metal and maintained at this temp. for a time varying from 1/2–6 min., then heated to a temp. above the upper crit. point of the metal, maintained at such temp. for 2–6 min., then cooled rapidly to a temp. below the lower crit. point of the metal, maintained at such temp. for 2 min. or less, reheated to a temp. near the lower crit. point of the metal, maintained at this temp. for 2–8 min. and then permitted to cool to atm. temp. This treatment gives the metal good drawing and stamping properties.

Protecting iron and steel from corrosion. A. FOLLIET and N. SAINDERICHIN. Brit. 337,562, Oct. 10, 1929. Articles of iron or steel are heated to 850–1050° in a mixt. comprising finely granulated Al free from Cu or Zn (the Al₂O₃ present as impurity not exceeding 6%), an inert material of high thermal cond. such as SiC free from graphite (this material constituting 40–60% of the volume of the mixt.), and 2–5% of one or more chlorides of non-ferrous metals alloyable with Fe, such as Cr, Ni or Mn. Double chlorides may be used and NH₄Cl 1.0–2.5% may also be added. A layer of an alloy of Fe, Al and the metal of the chloride used is formed on the articles.

Protecting iron and steel from corrosion. A. FOLLIET and N. SAINDERICHIN. Brit. 337,635, Nov. 28, 1929. A protective surface alloy is formed on iron or steel articles by embedding them in a mixt. of Al or Al-Mg alloy, oxides or oxidized ores of metals to be alloyed with the iron or steel (such as ilmenite, calcined pyrolusite, calcined garumite or roasted nickeliferous pyrites, chrome Fe ore or wolframite), and 2–10% of one or more chlorides of the same metals. By heating in mixts. of this kind surface alloys are formed such as chrome-Ni steel, Ti-Mn steel or the like. Cf. C. A. 24, 4254.

Open-hearth steel. ALEXANDER L. FIELD (to Republic Steel Corp.) U. S. 1,794,068, Feb. 24. A melted charge is refined in the presence of a slag in which a substantially constant percentage of Fe oxide is maintained by necessary tests and adjustments.

Open-bearth steel. GLENN E. HILLIARD. U. S. 1,793,442, Feb. 17. A molten bath of carbonized metal in the furnace is maintained in mobile condition and a strongly oxidizing flame is passed over the bath to effect oxidation of impurities. Finely divided scrap is introduced at such a rate of speed and in such volume that it is heated 'nearly to molten condition' while approaching the surface of the bath but is not excessively oxidized by the flame.

Hardening steel by nitridation. F. KRUFF A-G. Brit. 337,404, May 25, 1928. Nitridized articles of high resistance to stresses produced by high pressure or impact are made from steel free from Al and contg. Cr 1–20, C 0.25–3.0, W 0.5–3.0% and up to a total of 3% of Si, Mn, V and Mo (not more than 1% of any one of these). A

temp of about 500° and a time of about 70 hrs is suitable for the nitridation. *Cl. C A 24, 818.*

Treating steel scrap in open-hearth furnaces. *GLENN E HILLIARD U. S. 1,783,441, Feb 17* Fine metal scrap is introduced into open hearth furnaces in the form of a free falling stream of such volume and at such rate that the occluded air will be consumed by the furnace gases while the scrap is moving into charged position.

Stainless steel. *POLDINGEN A-G Ger 514,914, Sept. 2, 1923* An acid- and rust proof steel contains Ni 20-25, C 0.2-0.8, Cr 9-27 and not more than 0.5% Mn or Si. *Cl. C A 25, 1792*

Alloys resistant to corrosion at high temperatures. *R A HADFIELD Brit. 337,893, Sept. 11, 1929* An alloy suitable for elec. furnace resistors contains Fe together with Cr 24-35, W 3-5 and Al 3-7%, the proportions of C, Mn and Si being low (suitably about 0.6, 0.3 and 0.6%, resp.)

Aluminum alloys. *ALUMINIUM INDUSTRIE A-G Brit. 337,822, Oct. 12, 1928.* Elongation and bending and rolling capability of sheets, bands, wires, etc., of Al alloys of high tensile strength is improved by plating them with an Al alloy of lower tensile strength but of high elongation and good bending and rolling capability such as one contg Mn 1-4, Mg 0.5-6% and up to 1% Sb, or one contg Mn 0.5-2, Mg 0.3-1 and Si 0.5-2%, or up to 2% of Pb. Various details of manu. are given.

Aluminum alloys. *BIRMINGHAM ALUMINIUM CASTING (1903) Co., Ltd., and P. PRITCHARD Brit. 337,558, Oct. 0, 1929* Alloys comprising Al 94.5-96, Mg 3.5-5.0 and Mo 0.5% are made by melting Al and Mg and adding a melted Al-Mn alloy.

Aluminum alloys. *ALUMINIUM INDUSTRIE A-G Swiss 142,629, Sept. 23, 1929* An alloy suitable for vehicle buffers contains Al, with Mg 0.3-1, Si 0.5-2, Mo 0.5-2 and Cu 3.5-5%.

Copper alloys. *HANS KANZ. Swiss 141,610, Feb. 1, 1929* An alloy for grinding contains Zn 15-50%, up to 8% Sn, and Cu, with addn of any of the metals Al, Fe, Mn, Ni, Sb, Pb or Si. The example mentions an alloy contg Cu 62, Sn 4, Zn 30, Pb 2.5 and Ni 1.5%. *Cl. C A 25, 908*

Copper alloys containing also silicon, nickel, iron and zinc. *EUGEN VAPSEA. U S 1,792,944, Feb. 17* Casting alloys suitable for tools are formed mainly of Cu together with Si 1-4, Ni 1-5, Fe 4-6, and Zn 8-20%. *Cl. C A 24, 4503*

Iron alloys. *WILLOUGHBY S SMITH, HENRY J GARNETT and JOHN A HOLDEN Ger 512,727, April 23, 1928* An alloy of high permeability for use in elec. circuits contains 4-10% Cr, 1-4% Al, 0.05% C and the rest Fe.

Iron alloys. *VEREINIGTE STAHLWERKE A-G Brit. 338,071, Jan. 30, 1929* In the manu. of Fe alloys, the rusting tendency of the alloy is detd. rapidly by first testing in the normal way the power of the material to withstand vibration stresses, and then ascertaining the power of the material to withstand vibration stresses while exposed to the corrosive action of water or moist air during the vibration test (the method being based on the fact that the rusting of steel proceeds in the same proportion as the resistance to fatigue stresses is reduced in the corrosion fatigue test).

"Rustless" iron alloys. *FREDERICK M BECKET and JAMES H CATTICHERT (to Electro Metallurgical Co.) U S 1,793,153, Feb. 17* In prepg a low-C alloy of Fe and Cr of "rustless" character, the metal bath is side-blown with a blast contg 50% or more free O to effect oxidation of constituents of the bath including C and a minor proportion only of the Cr and the temp of the bath is raised to and maintained above 1600° solely by the exothermic reactions induced by the blowing operation, and the blowing is continued until the C content of the bath is reduced to 0.2% or less.

Iron-chromium alloys. *SKODA WORKS, PLESEN Brit. 337,767, April 18, 1929* Fe-Cr alloys contg Cr 20-40 and C 0.2-3% contain also Ti 1-15 (preferably 1-5) and up to 15% of Al (preferably 5% or less).

Spongy metals. *HENNING G FLOHM. U S 1,792,532, Feb. 17. See Fr 662,960 (C. A 24, 330)*

Steel alloys containing chromium. *F KRUPP A G Brit. 337,349, June 26, 1929* Austenitic Cr steel or Cr-Ni steel alloys contain also such a proportion of V or Ti or both as to enter into stable combination with the C dissolved by the austenitic basic mass and retain the alloy in austenitic form even when it is heated to 500-800°. The quantity of Ti or V is preferably at least twice that of the C.

Alloy steels. *W. MATHESIUS and H MATHESIUS Brit. 337,715, March 16, 1929* In making Ti steel, previously decarburized and deoxidized liquid steel is passed from a ladle through a discharge brick in the bottom into a second ladle contg pulverized or briquetted Ti thermic mixt.

Alloy steels. *R. SARGSON, Brit. 338,315, Oct. 18, 1929* Alloy steel articles

which have been nitrided at about 400-500° for about 5-100 hrs. are reheated to between 500° and the lower crit. point of the core material (preferably about 650°) to decrease the brittleness and increase the toughness of the case. An example is given of the treatment as applied to a steel contg C 0.32, Mn 0.56, Si 0.18, Cr 1.43, Mo 0.18 and Al 1.3%.

Alloy steels. J. T. WHITELEY (to I. W. Heyman and S. L. Zayon, trading as Multi-Steel Co.) Brit. 337,919 Jan. 24, 1929. Alloys are formed contg C 0.2-0.5% together with Mo 0.4-0.75, Cr 0.1-0.3, Si 0.2-0.5, Mn 0.5-1.0, Zr 0.5-1.5, Ti 0.1-0.3, Cu 0.5-0.9 and Ni 0.1-0.35%, and may be hardened by heating to about 925-950° and quenching.

Magnesium alloys. I. G. FARBERND A-G (Adolf Beck and Hugo Dibelka, inventors) Ger. 517,102 July 30, 1928. Addn. to 493,827 (C. A. 24, 2713). The method for rendering the surface of Mg alloys non-corrosive, described in 493,827, is modified by moistening the surface with dichromate soln. and drying, before carrying out the dichromate and HNO₃ treatment.

Magnesium alloy suitable for pistons of internal-combustion engines. WALTHER SCHMIDT (to I. G. Farbernd A-G.) U. S. 1,793,023, Feb. 17. Mg is alloyed with other metals such as Cu 15, Al 2 and Si 3%.

Strengthening magnesium and its alloys. I. G. FARBERND A-G. Brit. 337,706, March 11, 1929. The strength especially the compression yield point, of Mg and its alloys is improved by hot working (suitably at 200-500°).

Silver alloys. DEUTSCHE GOLD- UND SILBER-SCHNEIDANSTALT FORM ROESSLER. Ger. 514,772, Apr. 7, 1929. Alloys of Ag-Cu and Ag-Cu-Cd are improved by heating to glowing to a temp. just below that at which fusion commences and quenching. The process may be repeated with cold mech. working between the processes. In an example, a Cu-Ag alloy is heated to 730° and cooled in water.

Tungsten-carbide alloys. JOSEPH T. TERRI. U. S. 1,792,943, Feb. 17. For controlling the C content of a W carbide alloy, to prevent the formation of graphitic carbon, CaO is added to W and the mixt. is fused in a carbon crucible.

Steam and vacuum treatment system for drying spools of water-quenched wire. CHARLES B. CARR (to Western Elec. Co.) U. S. 1,793,275, Feb. 17. App. and various details of steam and vacuum treatments are described.

Pickling metals with acid solutions. VANDERVEER VOORHEES (to Standard Oil Co. of Ind.) U. S. 1,793,146, Feb. 17. An acid soln. such as a 5-10% soln. of H₂SO₄, for use in pickling metals such as iron or steel, is prepd. with an addn. of org. substances recovered from sludge derived from a cracked petroleum distillate sol. in dil. H₂SO₄, but free from sludge substances sol. only in strong H₂SO₄. These substances serve to inhibit corrosion of the metal.

Pickling bath for steel plates, wire, etc. EDWIN C. WRIGHT (to National Tube Co.). U. S. 1,792,958, Feb. 17. A 5% H₂SO₄ soln. is used with an addn. of about 0.05% of a sulfonated product of coal tar distillates b. 270-350° (other than anthracene oil), which serves as an inhibitor.

Purifying acid waste liquors such as those from pickling steel. JOHN T. TRAVERS (to Ohio Sanitary Engineering Corp.) U. S. 1,793,342, Feb. 17. Acid waste liquors contg. Fe salts are treated with a mixt. of lime and a solid waste product recovered from an alkali plant, the latter being in a quantity about 4 times that of the lime.

Case-hardening (nitriding) metal articles. JOHN J. EGAN (to Electro Metallurgical Co.) U. S. 1,793,309, Feb. 17. Articles such as iron alloys are heated in the presence of a nitriding agent such as NH₃ and at least one of the N oxides.

Local case-hardening of metals. AUBERT & DUTAL FRÈRES. Ger. 517,077, Nov. 18, 1927. The parts to be protected from nitridation are coated with a salt or a salt mixt., or with a glaze, enamel or the like, e. g., with KCl or with a mixt. of BaCl₂ 2, KCl 1.5, CaCl₂ 1 and NaCl 1.25 parts, or with a mixt. of PbO 7, sand 2 and borax 1 part. The parts to be protected may be painted with water glass and the powd. salt, etc., then applied.

Sealing metals to insulating materials such as glass. N.-V. PHILIPS' GLOEDLAMP-FABRIEKEN. Brit. 337,491, Aug. 12, 1929. In forming vacuum tubes, etc., comprising metal members such as rings or disks sealed to glass or the like, the metal at the point of union consists of an alloy preferably contg. 50-90% Fe together with one or more of the elements W, Mo, Ta, Nb, Co or Si, the coeff. of expansion of the metal being about the same as that of the insulating material with which it is to be united. Substances such as Cr, Al or Mg (preferably less than 5%) which form a coherent oxide film also may be added. Various structural details of manuf., etc., are described.

Apparatus for spraying metallic materials. THEO E. BRIDGER. Swiss 141,331,

Mar 21 1929 The metal is finely divided, centrifuged and subjected to an intermittent elec. current in an atm. of indifferent gas. Cf C A 24, 51.

Coating articles with vaporized or atomized metals, etc. KURT RICHTER. Ger 517,419 Dec 8, 1928. Low pressure vaporization or cathodic atomization of the coating materials is effected in a flattened vessel having a relatively small vapor space through which a gas stream is led at a high velocity. Back-diffusion of the coating material is thus diminished.

Protective coatings on metals. A. ASHBY Brit. 338,221, Aug 20, 1929 Metals such as iron or steel are provided with a protective coating or surface layer by treatment with a soln. of a chloride of As, Sb or Bi in a substantially non aq. solvent such as C_6H_6 , kerosene, acetone, $AmOAc$, $PtOAc$ or their mixts. or with the vapor of such a soln. if a volatile chloride is used. The material may be preliminarily coated with Zn if desired and a surface thus prep'd is suitable for further coating with paint, etc.

Coating and coloring aluminum alloys. A. PACE Brit. 338,204, Aug 16, 1929 A colored protective coating is formed by treatment with a soln. contg. chromates or dichromates or both in quantity not exceeding 0.17% by wt. calcd. as CrO_3 and a comparatively large quantity of an alk. substance such as not less than 1.5% of Na_2CO_3 . Various colors are produced depending on the compn. of the Al alloys, which may contain Sn, Zn, Mn, Sb, Mo, Cu, Ag, Sn or W, small quantities of metal salts may be added to modify the coloring.

Coating iron or steel sheets, etc., with metals such as zinc-aluminum alloy. AMERICAN ROLLING MILL CO. Brit. 357,300, July 29, 1929 The metal to be coated is passed successively through a flux, a bath of molten Pb and a bath of molten Zn-Al alloy contg. 3-30% Al. App. and various details of operation are described. Cf C A 25, 909.

Apparatus (with squeezing rolls) for coating metal sheets with materials such as tin. EDMUND W. SMITH and ROBERT H. SCOTT U. S. 1,794,093, Feb. 24. Structural features.

Determining the thickness of metal coatings such as tin and chromium. WESLEY A. RICHARDS and PAUL H. KRAMER U. S. 1,792,606, Feb. 17. The coating is liberated by a stripping acid such as 37% HCl so that the salt of the coating will be in reduced condition. The reduced salt is measured by means of 1 soln. and starch as an indicator in the reaction reagent.

Cement mixture suitable for using metals. WM. R. CHAPIN. U. S. 1,793,156, Feb. 17. Silico-Mn 65 and powd. borax glass 1-10 are used with fine steel filings to make up about 100 parts total. U. S. 1,793,157 specifies silico-Mn 40, powd. Cu 25, powd. borax glass 1-10 and fine steel filings to make a total of about 100 parts. Cf C A 24, 5281.

Using iron with other materials such as tin or nickel. WM. H. SMITH U. S. 1,793,757, Feb. 24. In forming alloys of Fe, finely divided iron such as "sponge iron" is assoc'd. with the other alloy elements such as Sn or Ni while still comparatively cold, and the assoc'd. materials are pressed into the general desired form so that they hold together by adhesion. They are then heated to a temp. somewhat below the m. p. of the iron to partly fuse the mass, heated up to alloying temp., and then subjected to pressure to effect final alloying and shaping.

Hard metal composition for cutting tools and dies. FLOYD C. KELLEY (to General Elec. Co.) U. S. 1,794,300, Feb. 24. A powd. mixt. consisting largely of W but contg. appreciable quantities of C and Co is treated with a small percentage of an alkyl resin and moistened with a solvent such as acetone. The material is pressed, heated sufficiently to vaporize the solvent and resin, further heated to sintering temp. and subjected to pressing while highly heated.

Use of tungsten carbide for making hard tough metal products. SAMUEL L. HOYT (to General Elec. Co.) U. S. 1,794,229, Feb. 24. Powd. W carbide is mixed with a cementing metal such as Co and the mixt. is pressed into the desired form to make tool bits or other articles, sintered into a hard mass and then subjected to pressure while subjected to a sintering temp. Cf C A 24, 5016.

Apparatus for flexing metal wire or strips such as iron or steel while immersed in a liquid such as sulfuric acid to remove surface impurities. FRANK M. DORSEY (to Madsenell Corp.) U. S. 1,793,914, Feb. 24.

Filaments of material such as metals. GEORGE F. TAYLOR (to Baker & Co.) U. S. 1,793,529, Feb. 24. Material such as a metal of low or high m. p. sheathed with borax, glass or quartz is drawn into a cored filament while heated to render it sufficiently fluent and the cored filament is solidified close enough to the heated substance so that the substance will flow to the point of solidifying. App. is described.

Indented metal foil. ALUMINIUMWERKE A.-G., RORSCHACH. Brit. 337,716, May 24, 1929. Metal foils with deep indentations are formed by passing smooth foil through gossamer rollers together with paper which is subsequently burned away in an annealing process (initially at 300–550°).

Metal bearings. HENRY L. HARTON (to General Motors Research Corp.). U. S. 1,702,565, Feb. 17. Powder bearing metal such as finely comminuted Babbitt alloy is placed in a bearing back and subjected to pressure to cause it to unite with the back in good heat conducting relation.

Magnetic material comprising nickel, chromium, cobalt and iron. GUSTAV W. FLUMEN (to Bell Telephone Laboratories Inc.). U. S. 1,702,481, Feb. 17. A material comprising Ni 10–80 and Co 5–80, Fe between about 60% of the entire Ni-Fe-Co content and Cr up to 12% is heat treated to give it good magnetic properties with small magnetizing forces, and is suitable for use in making loading coils of signaling systems, etc. Cf. C. A. 25, 678.

Metal for use in spark-plug manufacture. A. A. POSTNIKOFF and R. P. ZERKINDEN. Brit. 337,890, Sept. 7, 1929. The metal surfaces exposed to the combustion gases are coated with Cr or a Cr alloy to a thickness of 0.0001–0.00025 in.

Cutting elements for paper-making engines. CLARENCE W. NOBLE (to Noble and Wood Machine Co.). U. S. 1,702,516, Feb. 17. The side portions of cutting edges are formed of relatively hard metal such as Cr and the central portions of softer metal such as steel.

Welding electrode. JOHN G. RITTER (to Westinghouse Elec. & Mfg. Co.). U. S. 1,701,983, March 3. Welding electrodes formed of an alloy of Fe 65–95%, minor impurities and Ni 35–55% produce welded joints substantially free from residual stresses by deposition at a temp. near or below its low annealing temp. Various details of temps. for working with varying comps. of alloy are described.

Are electrodes for welding. J. H. ATTCHISON, H. BULL and L. JOHNSON. Brit. 337,549, Oct. 25, 1929. Billets are formed of one or more metals of an alloy to be welded (such as Cr steel for welding "stainless steel") and are provided with one or more longitudinal cores or coverings of the other metal or metals of the alloy to be welded (such as Ni), and the composite billet is worked into composite electrodes by rolling, forging or drawing. Various details and modifications of manuf. and comps. for electrodes are given.

Electrode for metal deposition by the electric arc process. ARTHUR P. SROOG-MANRGA (to Quasi Arc, Inc.). U. S. 1,701,012, Feb. 21. An electrode is formed with a flux-forming covering and a ferrous core contg. a quantity of Mn dependent on its diam. and varying inversely as the diam. from 2.25% to 4.0% for cross sections of electrode cores from 0.012 to 0.005 sq. in. This proportion of Mn avoids a brittle zone in the deposit. Cf. C. A. 24, 52.

10—ORGANIC CHEMISTRY

CHAS. A. ROUILLER and CLARENCE J. WYATT

The determination of the empirical formula of a hydrocarbon. EDWARD W. WASHBURN. *Bur. Standards J. Research* 5, 897–900 (1930).—The precision aspects are discussed of the problem of detg. the mol. wt. and H content of a hydrocarbon, and of combining the results so as to obtain the empirical formula. "A detn. of the H (or other) aliph. no. may, in some instances, be substituted for the mol.-wt. detn. or for the combustion analysis, or may be utilized to decrease the accuracy which would otherwise be required in either or both of these detns." By following the lab. procedure outlined it should be possible to det. the empirical formula of any pure hydrocarbon contg. not more than 100 C atoms.

H. B. SCHMIDT

Formation of carbon tetrachloride from the elements. ALFRED STOCK, HERMANN LUX and WERNER WUSTROW. *Z. anorg. allgem. Chem.* 195, 149–57 (1931).—Specially prepd. C made from filter paper and well outgassed was used. This prepn. was placed in a quartz tube in a furnace at either 400° or 500° (only 2 temps. were studied) and measured amts. of Cl₂ were admitted, the contents of the tube were analyzed at various periods. CCl₄, HCl and traces of CO₂ were found. Heating periods were for as long as 90 hrs. Activation of the C was attempted. The effect of HCl on the reaction vessel was also studied; in 70 hrs. 37 cc. HCl was quantitatively converted into H₂O. Cl₂ had practically no effect on the glass.

WILLIAM H. VAUGHAN

Pentano. C. R. NOLLE. *Org. Syntheses* XI, 84–C (1931).—Details are given for

the prepn of $C_{11}H_{18}$ from $Me(CH_2)_9CHBrMe$ through the Grignard reaction; yield, 50-37% C. J. WEST

Carotene. III. Hydrogenation and optical properties of carotene and its hydrogenated derivatives. JAMES H. C. SMITH *J Biol Chem* 90, 607-605 (1931), cf. *C. A.* 24, 3794 —By reduction of carotene (I) in Et_2O with $Al-Hg$, filtering and drying *in vacuo*, dihydrocarotene, $C_{40}H_{58}$, was obtained as a noncrystallizable orange compd. $[\alpha]_D^{25} 41.9^\circ$ (Et_2O), 38.0° ($EtOH$). I had $[\alpha]_D^{25} -63^\circ = 17\%$ (CS_2). The degree of hydrogenation by Adams' catalyst varied with the amt of catalyst. With 5.2 g I in cyclohexane and 0.126 g catalyst, 9 mols H_2 was absorbed to give a colorless oil, $C_{40}H_{70}$, octadecahydrocarotene (II), $[\alpha]_D^{25} 7.56^\circ$ (CS_2), $d_4^{25} 0.8828$, $d_4^{20} 0.8524$, $n_D^{25} 1.4962$ — 0.000367; M_p 180.3 (calcd for straight chain 178.2). With 0.85 g I and 1.84 g catalyst, 10.23 mols H_2 was absorbed to give a clear viscous oil, $C_{40}H_{78}$, eicosahydrocarotene (III), $b.p.$ 206°, $[\alpha]_D^{25} 0.337^\circ$, $n_D^{25} 1.4821$, $d_4^{25} 0.8748$, M_p 181.42 (calcd for straight chain 180.32). From the low mol. exaltations of II and III the presence of the cyclopropane rings is deduced. The steady decrease in rotation on hydrogenation shows that the 2 linkages hardest to sat are those producing asymmetry. The 2 bonds reducible by $Al-Hg$ are probably attached to rings (Kuhn and Winterstein, *C. A.* 23, 4682). A structure of 9 conjugated double bonds, conjugated further with 2 cyclopropane rings, is assigned to I, this agrees with the formula of Karrer, *et al* (*C. A.* 25, 519) K. V. THIMANN

A method for the preparation of substituted acetylenic hydrocarbons. RENE TACHET *Compt rend* 191, 854-6 (1930) —Methylacetylenes are satisfactorily prepd by treating a Na acetylide with Me_2SO_4 , $PhSO_2Me$ or $p-MeC_6H_4SO_2Me$. The reaction with aromatic SO_2H acids appears to be perfectly general, 45-60% yields resulting from treating $C\equiv CNa$ with various sulfonic esters for 3 hrs. at 80° , decomp. with H_2O , extg. with Et_2O , drying over $CaCl_2$, distg. *in vacuo* and redistg. at atm pressure. $C\equiv CNa$ (I) plus $p-MeC_6H_4SO_2Et$ (II) gave 45% of 3 nonine, b. $155-7^\circ$, $d_4^{25} 0.763$, $n_D^{25} 1.429$, $M. R.$ 41.9 (calcd 41.77). $C\equiv CNa$ (I) with II gave 3-decine (47% yield), b. $175-6^\circ$, $d_4^{25} 0.765$, $n_D^{25} 1.433$, $M. R.$ 46.6 (calcd 46.4). I with $p-MeC_6H_4SO_2Bu$ gave 5-hendecene (60% yield), b. $193-6^\circ$, $d_4^{25} 0.785$, $n_D^{25} 1.437$, $M. R.$ 52.0 (calcd 51.1). I with $PhSO_2C_6H_5$ gave amylallylacetylene, b. $74.6-5.5^\circ$, $d_4^{25} 0.788$, $n_D^{25} 1.445$, $M. R.$ 46.2 (calcd 46). By-products of these reactions may be ethers produced by heating the sulfonic esters, also from sapon of the esters or 1-chloroacetylenic hydrocarbons resulting from the presence of small quantities of sulfonyl chloride. These chloroacetylenes are especially hard to remove so care should be taken to use esters contg. none of the chloride. C. H. P.

Preparation of chlorinated acetylene derivatives. STEPHAN LANGGUTH *Chimie & industrie* 25, 22 5 (1931) —Directions are given for the prepn of $SiCl_3C\equiv CCl$, $SiCl_3C\equiv CCl_2$, $C_6H_5C\equiv CCl$, $CCl_3C\equiv CCl$ and $(PhN)C\equiv CCl$. A. PAPIEUX COUTURE

The higher alcohols of the paraffin series and their technical significance. SCHRAUTH *Chem. Ztg* 55, 3, 17 8 (1931) —The higher alcs are now available, made by reduction of the corresponding acids. Various uses are indicated, chiefly as the sulfonic and sulfonic acid deriva. in medicine and as emulsifying agents, as esters in perfumery and as waxes T. H. CHILTON

Some α,α -dichloro and α,α -diethoxy tertiary alcohols and the hydroxy aldehydes derived from them. A. VVY *Bull soc chim* [4], 49, 12-8 (1931) —The general problem of prepn. α -HO aldehydes of the type $R_2C(OH)CHO$ (I) from $R_2C(OH)CHCl_2$ (II) and $R_2C(OH)CH(OEt)_2$ (III) was investigated. Type II compds were prepd by treating Cl_2CHCO_2Et (IV) with the suitable organomagnesium bromide since the bromides gave better yields than the iodides used heretofore. IV was obtained in 90% yield by boiling for 6 hrs. 100 g Cl_2CHCO_2H with 300 g abs. alc. in the presence of 3 cc. H_2SO_4 . $Me_2C(OH)CHCl_2$ obtained in 74% yield, m. 8° , b. 38° , $n_D^{25} 1.2507$, $d_4^{25} 1.2363$, $n_D^{25} 1.4598$. $Ft_2C(OH)CHCl_2$ (V), obtained in 70% yield, b. 76° , $d_4^{25} 1.1999$, $d_4^{25} 1.1750$, $n_D^{25} 1.47102$. $EtCOCHCl_2$ is obtained in the proportion of 12 g to 100 g V, in the prepn. of V. $Pr_2C(OH)CHCl_2$ and $Bu_2C(OH)CHCl_2$ were obtained by this method in very poor yield (15%) and in an impure state. $Ph_2C(OH)CHCl_2$ (VI), obtained as chief product along with Ph_2 and an unidentified very slightly sol. compd, b. 200° , crystals from a mixt. of Et_2O and petr. ether, m. $95-6^\circ$. $PhCH_2AlBr$ acted on IV to form dibenzyl as the only isolated product. Type III compds were prepd more readily and in better yield than type II compds by the action of the suitable organomagnesium compd on $(EtO)_2CHCO_2Et$. As the reaction is violent cooling is necessary. Type III compds are more sensitive to the action of mineral acids than type II compds. $Me_2C(OH)CH(OEt)_2$ (VII), obtained in 70% yield, b. 75° , $d_4^{25} 0.9478$, $d_4^{25} 0.9277$, $n_D^{25} 1.41109$. $Et_2C(OH)CH(OEt)_2$ (VIII), obtained in 83% yield, b. 95° , $d_4^{25} 0.9400$, d_4^{25}

0.9230, n_D^{25} 1.4235. $\text{Pr}_2\text{C}(\text{OH})\text{CH}(\text{OEt})_2$ (IX), obtained in 70% yield, b_p 98°, d_4^{20} 0.9261, d_4^{25} 0.9073, n_D^{25} 1.4238. $\text{Bu}_2\text{C}(\text{OH})\text{CH}(\text{OEt})_2$ (X), obtained in 70% yield, b_p 101°, b_r 122°, d_4^{20} 0.9167, d_4^{25} 0.9001, n_D^{25} 1.4315. $\text{Ph}_2\text{C}(\text{OH})\text{CH}(\text{OEt})_2$ (XI), obtained in 55% yield, b_p 170°, m 125°. Attempts to prep. type I compds. by heating type II compds. with CaCO_3 suspended in water, Na_2CO_3 , PbO and water for various periods at various temps. never resulted in the isolation of type I compds. though small amounts of impure $\text{Pr}_2\text{C}(\text{OH})\text{CHO}$ (XII) were obtained by refluxing 72 g. V for 36 hrs. with 300 cc. water and 50 g. CaCO_3 extg. with ether, drying and fractionating, when 15 g. of a yellow oil b_p 55–57° was obtained which gave a semicarbazone (XIII) of XII, m 204–5°. Hydrolysis of XIII with aq. $(\text{CO}_2\text{H})_2$ gave a liquid b_p 28°, semicarbazone in 210°, oxime in 62°. Analysis indicates these compds. are derived from XII after elimination of 1 mol. of water. Hydrolysis of type III compds. with warm dil. HCl yielded type I compds., but with cold concd. HCl intermediate compds. of type XIV apparently resulted according to the scheme $2\text{R}_2\text{C}(\text{OH})\text{CH}(\text{OEt})_2 \rightarrow \text{R}_2\text{C}(\text{OH})\text{CH}(\text{OH})\text{OEt} + \text{R}_2\text{C}(\text{OH})\text{CHO} \rightarrow \text{O}=\text{CR}_2\text{CH}(\text{OH})\text{O}=\text{CHC}(\text{OEt})_2\text{(XIV)} + \text{H}_2\text{O}$. Type XIV compds. on

further hydrolysis yield type I compds. When 10 g. of VII, 2.5 g. water and 3.5 g. concd. HCl stood 24 hrs. in a flask and were then extd. with ether there was obtained 5 g. of a sticky liquid, b_p 105°, whose properties and analysis correspond to those predicted for the corresponding type XIV compd. Warm dil. HCl on VII gave a high boiling material not identified. Hydrolysis of 9 g. VIII with cold dil. HCl gave 5 g. liquid, b_p 114–5°, which did not give a semicarbazone and has strong reducing properties and is probably a polymer of $\text{Et}_2\text{C}(\text{OH})\text{CHO}$ (XV). Warm dil. HCl on VIII gives 65% of impure XV, characterized by its semicarbazone as a type II compd. Hydrolysis of 7 g. IX with cold HCl gave 3 g. of an apparently type XIV compd., crystals from ether, m 87°. Hydrolysis of X with cold HCl apparently produces (from analysis and mol. wt. detn.) a type XIV compd. (XVI), m 90–1°. XVI (0.9 g.), distil. with 20 cc. water and 2 cc. HCl, produces 0.7 g. oil, b_p 94–7°, d_4^{20} 0.9020, d_4^{25} 0.9040, n_D^{25} 1.4379, whose mol. wt., analysis and analysis of its semicarbazone, m 141–2°, prove it to be pure $\text{Bu}_2\text{C}(\text{OH})\text{CHO}$.

OREN E. SHEPPARD

The use of calcium chloride in the dehydration of alcohol. K. B. EDWARDS AND R. LACEY. *J. Soc. Chem. Ind.* 49, 422T (1930).—Alc. residues are concd. to 80% by distn. Then 20 g. of fused CaCl_2 per 100 cc. alc. is added. Distn. yields 75% of 93.5% alc. and 5% of "tail." The alc. remaining with CaCl_2 in the still is recovered by adding water and distg. Simple tests for the concn. of the alc. have been worked out; they depend on the miscibility of one drop of alc. in 5 cc. of various solvents. Kerosene, miscible when the concn. is 99.35%; white spirit, 98.35%; com. xylene, 97.45%; permanent cloudiness with 3 drops of 97.25%. The tests are to be carried out between 15° and 18°.

ALBERT L. HENNE

Triethylcarbinol. W. W. MOYER AND C. S. MARVEL. *Org. Syntheses* XI, 98–100 (1931).—Detailed directions are given for the prepn. of Et_3COH from EtMgBr and $(\text{EtO})_2\text{CO}$; the yield is 82–85%. The following were prepd. by the use of the proper Grignard reagent and $(\text{EtO})_2\text{CO}$: Pr_2COH 73, Bu_2COH 84, Am_2COH 75 and $(\text{C}_6\text{H}_5)_2\text{COH}$ 72%.

C. J. WEST

Aliphatic ethers. HERBERT HEYSTOCK. *J. Chem. Soc.* 1931, 371–2.— PrI and iso- BuONa , shaken 12 hrs. at 10° and then heated 1 hr. at 100°, give 67.2% of Pr is-Bu ether , b_p 106°, d_4^{20} 0.7549, n_D^{25} 1.3832. iso- Pr Bu ether , b_p 108°, d_4^{20} 0.7504, n_D^{25} 1.3889 (72.4% yield). iso- BuOPr , BuOPr , iso- PrOPr , EtOPr , MeOPr , iso- BuOBu , iso- PrOBu , EtOBu , MeOBu , iso- PrOEt , iso- PrOMe , iso- BuOEt and iso- BuOMe , heated with Na K in N at 190–200° for 24–100 hrs., are unchanged, it thus appears that in ethers the bonds between the simple alkyl radicals and the O atom are stronger than the corresponding bonds holding heavier radicals contg. highly branched chains.

C. J. WEST

Syntheses with β, δ' -dichlorodiethyl ether. II. Heterocyclic compounds containing two members of the oxygen group in the ring. 1,4-Selenoxane and its derivatives. CHARLES S. GIBSON AND JOHN D. A. JOHNSON. *J. Chem. Soc.* 1930, 266–72. *Chemistry & Industry* 49, 826 (1930); *cf. C. A.* 25, 661— $(\text{ClCH}_2\text{CH}_2)_2\text{O}$ or $(\text{ClCH}_2\text{CH}_2)_2\text{S}$ and Na_2Se in water, heated 4–5 hrs. while II is passed through the mixt. give 1,4-selenoxane (I), purified through the dibromide which is obtained in 33–40% yields; 164 g. dibromide gives 60 g. pure I; b_p 160°, b_r 69.5°, b_m 167.5–85°, m –21.5°, n_D^{25} 1.5480, d_4^{25} 1.575, d_4^{20} = 1.597 – 0.00135t; γ^{25} 42.08, γ^{20} 40.80, γ^{15} 39.75, γ^{10} 38.67 dynes/cm.; I is a colorless liquid with penetrating but not unpleasant odor, it is volatile in CCl_4 vapors;

SnCl_4 gives a cryst compd decompd by atm moisture; I and HNO_3 (d. 1.42) give a deep red soln from which the hydroxynitrate sepd on cooling, it decomp 140-1° and detonates on heating in a sealed tube over a small flame, its aq soln gives the dibromide with HBr in the cold. I and Cl in CCl_4 give the dichloride, m 127-9° (decompn), Br in CCl_4 gives the dibromide, yellow, m 132° (decompn); the aq soln gives no ppt on cooling but HBr ppts the dibromide, diiodide, violet, m 106-7°. I and MeI in abs EtOH for 5 hrs. or I and MeI for 2 hrs give the methiodide, m 171°. I and HgCl_2 in EtOH give the mercurichloride m 179°, chloroaurate, golden yellow, decomp 142-4°. chloroplatinate orange yellow m 149° (decompn). The dibromide and most Ag_2O give a highly hygroscopic oil which gives the hydroxynitrate with HNO_3 . The dibromide and dry NH_3 in C_6H_6 give the compd $\text{C}_{11}\text{H}_{11}\text{ON}_2\text{Br}_2\text{Se}$, begins to decomp 93° and turns brownish yellow 130°. It is decompd by water, similar compds were formed with pyridine, piperidine and H_2N_2 . The coordination compd of the dibromide with $\text{C}_{11}\text{H}_{11}(\text{N}_2)_2$ yielded $\text{C}_{11}\text{H}_{11}(\text{N}_2)_2$, HBr on treatment with aq LiOH . The parachor, 245.2, confirms the value of 62.5 for the at parachor of Se . C. J. WEST

Formation of sulfonium chlorides and of unsaturated substances by the action of water and of aqueous alcoholic potash on β, β' -dichlorodiethyl sulfide. JONV S. H. DAVIES and ALBERT E. OXFORD. *J. Chem. Soc.* 1931, 224-30.—The hydrolysis of $\text{S}(\text{C}_2\text{H}_4\text{Cl})_2$ (I) is generally assumed to occur in 2 stages: $\text{I} + \text{H}_2\text{O} \rightarrow \text{ClC}_2\text{H}_4\text{SC}_2\text{H}_4\text{OH}$ (II) + HCl , $\text{II} + \text{H}_2\text{O} \rightarrow \text{S}(\text{C}_2\text{H}_4\text{OH})_2$ (III) + HCl . It is probable that II has not been isolated, attempts to prep II from $\text{HOCC}_2\text{H}_4\text{SH}$ and $\text{ClC}_2\text{H}_4\text{Cl}$ or from $\text{ClC}_2\text{H}_4\text{SH}$ and $(\text{C}_2\text{H}_5)_2\text{S}$ failed. II, evapd in the min quantity of water at a low temp under diminished pressure gives a sirup, invol in CHCl_3 and Me_2CO and hence not II or III, but readily sol in cold H_2O to give a soln contg much Cl ion. I yields a diphenylurea, m 128.5-9.5°, the methiodide is a viscous red oil, the *p*-toluenesulfimine formed from chloramine-T, m 86-7°. III and $\text{HOCC}_2\text{H}_4\text{Cl}$ heated on the water bath for 13 hrs, give *tri- β -hydroxyethylsulfonium chloride* ($\text{HOCC}_2\text{H}_4\text{S}^+\text{Cl}^-$), m 125-6°, the yield is largest with 3 mols $\text{ClC}_2\text{H}_4\text{OH}$, it develops no acidity when dissolved in water and the Cl content can be acid volumetrically, it is sol in concd HCl but no I is formed on prolonged boiling. III and I, heated 6 hrs at 65-75°, give *sulfidobis- β -hydroxydiethyl sulfide* 1,3-di- β -hydroxyethiochloride, $\text{Cl}(\text{HOCC}_2\text{H}_4)_2\text{SC}_2\text{H}_4\text{SC}_2\text{H}_4\text{S}(\text{C}_2\text{H}_4\text{OH})_2\text{Cl}$, m 101.5-3°, with HCl at 100° only I is formed, the chloroplatinate, yellow, amorphous, m 133-4°. This explains why II has not been isolated by the hydrolysis of I, for, even if it were formed momentarily, it might unite with itself and most certainly would combine with III to form sulfonium chlorides $\text{HOCC}_2\text{H}_4\text{SC}_2\text{H}_4\text{SC}_2\text{H}_4\text{OH}$ and $\text{ClC}_2\text{H}_4\text{OH}$, heated at 100° for 14 hrs, give 2 sulfonium chlorides, probably IV and *sulfidobis- β -hydroxydiethyl sulfide* 1,2-di- β -hydroxyethiochloride, m 120-15°. $\text{HOCC}_2\text{H}_4\text{SC}_2\text{H}_4\text{SC}_2\text{H}_4\text{OH}$ and $\text{HOCC}_2\text{H}_4\text{Cl}$ heated on the water bath for 8 hrs, or $\text{HOCC}_2\text{H}_4\text{Cl}$ and 1,4-dithiane, heated at 100° for 28 hrs, give 1,4-dithiane 1- β -hydroxyethiochloride (IV), m 175°, mercurichloride, m 95-6°. Passing dry HCl into III at 100-6° gives 56% I and 44% of *oxidobis- β -chlorodiethyl sulfide*, b₁ 152°, the latter and PhOH with EtONa give the *phenoxy deriv*, m 55-7°. Hydrolysis of I with 2.25 vols water gives IV, but this was not found when 6 vols water was used. The sirup obtained with 4 mols water, heated with concd HCl 40 min at 100°, gives I and a small quantity of $\text{C}_{11}\text{H}_{11}(\text{SC}_2\text{H}_4\text{Cl})_2$. If 10 vols HCl is used, I is the sole product. III and 2 mols HCl in 2.5 vols water, heated 5 hrs, give IV and other sulfonium chlorides. Hydrolysis of 1 mol I with 4 mols 20% LiOH KOH ($\text{EtOH H}_2\text{O} = 4:1$) gives a mixt of divinyl sulfide, b 85° (sulfimine, m 91-3°), β -ethoxyethyl vinyl sulfide, b₁ 65°, d₄²⁰ 0.9532, camphor like odor (mercurichloride, m 152-3°), $\text{S}(\text{C}_2\text{H}_4\text{OEt})_2$, b₁ 101-2° and some $\text{EtOC}_2\text{H}_4\text{SC}_2\text{H}_4\text{OH}$, b₁ 117.5°. With 1 mol EtOH KOH , there was isolated β -chloroethyl vinyl sulfide, b₁ 71-2° (sulfimine, m 101.5-3°), mercurichloride, amorphous, decomp about 150°, with HCl this yields α, β' -dichlorodiethyl sulfide, b₁ 68-9°. C. J. WEST

Heptylamine W. H. LYCAN, S. V. PUNTAIBEKER and C. S. MARVEL. *Org. Syntheses* XI, 58-60 (1931).—Reduction of $\text{Me}(\text{CH}_2)_6\text{CH NOH}$ with EtOH and Na gives 60-73% of $\text{Me}(\text{CH}_2)_6\text{NH}_2$, b 152-7°. By use of the same method the following may be obtained: BuNH_2 , in 50-60% yields, *sec-BuNH}_2*, and cyclohexylamine. C. J. WEST

Acrolein acetal. E. J. WITZEMANN, WM. LLOYD EVANS, HENRY HASS and E. F. SCHROEDER. *Org. Syntheses* XI, 1-2 (1931).— $\text{ClCH}_2\text{CH}_2\text{CH}(\text{OEt})_2$ (167 g) and 340 g powd KOH heated at 210-20°, give 75% of $\text{CH}_3\text{CH}(\text{OEt})_2$, b 122-6°. Fe retorts should be used when many runs are to be made. C. J. WEST

β -Chloropropionaldehyde acetal. E. J. WITZEMANN, WM. LLOYD EVANS, HENRY HASS and E. F. SCHROEDER. *Org. Syntheses* XI, 26-7 (1931).— EtOH HCl , satd at 0°

and treated with CH_3CHCHO at about 0° , gives 34% of $\text{ClCH}_2\text{CH}_2\text{CH}(\text{OEt})_2$, b. $58-62^\circ$. C. J. WEST

Heptaldoxime. E. W. BOUSQUET *Org. Syntheses* XI, 54-6 (1931) — $\text{Me}(\text{CH}_2)_5\text{CHO}$, $\text{NH}_2\text{OH} \cdot \text{HCl}$ and Na_2CO_3 give 81-93% of $\text{Me}(\text{CH}_2)_5\text{CHNOH}$. Cyclohexanone oxime is obtained in the same manner. C. J. WEST

Fatty acid distillation. A comparison of intermittent and continuous operation. OTTO KREBS *Teer u. Bitumen* 28, 421-4, 440-3 (1930) — Batch distn with superheated steam and continuous distn under reduced pressure (60 mm) are described with some detail and the economics of the 2 methods compared. Vacuum distn permits an easy sepn of $\text{C}_{17}\text{H}_{33}\text{CO}_2\text{H}$, b. 255° . $\text{C}_{17}\text{H}_{33}\text{CO}_2\text{H}$ boils only slightly lower than $\text{C}_{17}\text{H}_{33}\text{CO}_2\text{H}$, their sepn, therefore, requires a very efficient fractionation column. K. H. ENGEL

Addition of alkali alcoholates to acid esters. IV. Addition of sodium ethylate to formic ester. F. ANICKES (WITH A. WACHTEROWITZ, H. DU MONT AND O. LÜCKER) *Ber* 63B, 3012-27 (1930), cf. *C. A.* 21, 1628; Scheibler, *C. A.* 20, 2824, 21, 1795 — S. says in his 1st paper that NaOCOEt (I) is obtained quantitatively according to the equation $\text{HCO}_2\text{Et} + \text{NaOEt} = \text{HC(ONa)(OEt)}$, (II) = I + EtOH by letting NaOEt and 1 mol HCO_2Et stand in Et_2O 16 hrs, evap. *in vacuo*, adding another 0.5 mol HCO_2Et , letting stand again and once more evap. off the excess of HCO_2Et , along with the EtOH formed, *in vacuo*. A. showed that the product was not I (*C. A.* 21, 1628) but the conclusion drawn from analyses of the apparently homogeneous substance that it was quite pure II has now been found to have been incorrect. Contrary to S.'s statement, a great part of the HCO_2Et used is decompd. into CO and EtOH and a part of the EtONa is not added at all, so that the product consists in great part of NaOEt with more or less EtOH of crystn. (around 30%) and some HCO_2Na (8%) and can, therefore, contain only about 10-20% II. With such a product, the Zeisel EtO detn. can be carried out only with certain precautions in a modified app., the far too low results obtained in the usual Meyer app. had led to an erroneous interpretation, and, moreover, the apparent absorption of 1 mol HCO_2Et by the NaOEt (calcd. from the nmt. of HCO_2Et used and that recovered from the Et_2O and EtOH vapors drawn off) was merely due to the afore-said decompn. of the HCO_2Et . According to S.'s 2nd paper, I is prepd. by splitting EtOH off from the (assumedly almost pure) II at 40° and completing the reaction by repeated treatment with HCO_2Et and splitting off of EtOH . No proof (by detn. of the CO) is given of the decompn. $\text{I} + \text{H}_2\text{O} = \text{NaOH} + \text{CO} + \text{EtOH}$ which S. describes as characteristic but which A. has never observed. The apparent success of the new method of prepn. is due to the fact that the temp. is raised by the "vigorous reaction" which takes place; the decompn. of the HCO_2Et , which is slow at room temp., is now very rapid and the product is not I but NaOEt , more or less free of EtOH of crystn., depending on the length of the last heating *in vacuo*. In both methods of prepn. the HCO_2H present in the product (as the acid and as HCO_2Et) can be detd. by reduction of HgCl_2 or by back titration after sapon. with excess of hot alkali. The HCO_2H is present partly as HCO_2Na because ordinary alic. and HCO_2Et not dried over P_2O_5 are used in the process. This is also shown by analysis of the salt mixt. obtained by passing CO_2 into the Et_2O suspension of the supposed I; the HCO_2Na remains unchanged, the NaOEt is converted into EtOCO_2Na and a part of the total HCO_2H is recovered as HCO_2Et , presumably from II, for I could not regenerate HCO_2Et with CO_2 . A time consuming method, based on the analytical detn. of the total Na, total HCO_2H , the HCO_2Na (by CO_2 decompn.), total EtO (modified Zeisel method) and EtOH of crystn. (Zerevitinov), was worked out to decide whether the product contained I or II (later it was found that light could be shed on this question by other means), the results showed that the product prepd. by either of S.'s 2 methods cannot be anywhere nearly pure I; that prepd. by the 1st method may contain up to about 25% II and that prepd. by the 2nd method none at all, and neither can possibly contain a I with the properties given by S. With none of the many preps. was an evolution of CO on decompn. with H_2O observed. That the ester still held bound in the reaction product is almost certainly present as II was shown by prepg. II in pure form, when it is heated *in vacuo* at 40° according to S.'s 2nd method for prepg. his supposed I, soon only the formate- HCO_2H remains and the end product is a mixt. of EtOH -free NaOEt + HCO_2Na . When EtOH -free NaOEt in Et_2O suspension is allowed to react at 0° with HCO_2Et , the HCO_2Et does not decomp. into CO , and 1 mol. (not more) is bound, the addn. product, which is almost insol. in Et_2O , is an exceedingly hygroscopic microcryst. powder which decomps. into CO at room temp. Attempts to obtain from it derivs. of HC(OEt)_2 with ClCOEt or PhCH_2Br failed so that it could not be proved directly that it has the structure II. The method previously used to det. the structure of ester alcoholate addn. products (cleavage of

equiv. amts. of MeOH and EtOH with H_2O or of NaOMe and EtONa with CO_2 from the mixed Et ester methylate or Me ester-ethylate) does not have the significance then ascribed to it, for, contrary to expectation, the esters may undergo an exchange of the alkyl (Umschlingung) in H_2O under the influence of the mol. alcoholate. The addn. product dissolves when treated in Et_2O suspension with $(CO_2Et)_4$ and HCO_2Et is set free, correspondingly, the $(CO_2Et)_4$ addn. product does not react with HCO_2Et . C. A. R.

3-Methylpentanoic acid. E. B. VILLET, C. S. MARVEL AND C. M. HUYER. *Org. Syntheses* XI, 26-8, (1931) — $EtCHMeCH_2(CO_2Et)_2$ sapon. with KOH and the acid liberated with H_2SO_4 and refuzed 3 hrs., gives 62-5% of $EtMeCHCH_2CO_2H$ bp 193-6°. Caproic acid was obtained in 75% yield by this method. C. J. WEST

The amides of α -methylbutenoic acids. P. BATTALINI, L. ENOCH and M. DEKOKER. *Bull. sci. acad. sci. Belg* [5] 16, 721-4 (1930), (cf. *Mem. acad. sci. Belg* 9, 16 (1927) and C. A. 22, 2304). — The amides, acids and nitriles of α -methylbutenoic acids were prep. and certain phys. properties observed. The configuration of the angles and tie nitriles were detd. by comparison with crotonic and isocrotonic nitriles and with 3,4- and 2,6-Me- C_6H_3CN . The butenoic nitriles, amides and acids were prep. by the dehydration of the cyanohydrins with P_2O_5 or $SOCl_2$, preferably the latter. The methyl nitriles were converted into methylchloroacetanor nitriles by addn. of HCl. Treatment with quinoline gave the isomeric butenoic nitriles. The following compds. were prep. (I) α -methyl-3-chlorobutanor nitrile, bp 64-5°. (II) α -methyl-3-chloropropionitrile, bp 73-4°. (I) on treatment with quinoline gave 2 isomers of α -methylcrotonitrile, bp 121-2° and 137-8°. The high-boiling isomer (III) has been previously described. The low-boiling isomer (IV) gave d^{20}_D 0.8106, n^{20}_D 1.4204, n^{20}_D 1.4233, n^{20}_D 1.4387, E_m 0.64. The $ClCH_2CHCH_2CN$ on treatment with quinoline gave α -ethylacrylonitrile (V), bp 115.0-115.4, d^{20}_D 0.8159, n^{20}_D 1.4144, n^{20}_D 1.4175, n^{20}_D 1.4243, E_m 0.51. The amide obtained from III in 75-8°, that from IV in 127-8°, and that from V in 57-7-5°. The first 2 amides both yield tie acid on hydrolysis or treatment with HNO_3 . Also in *Bull. sci. acad. sci. Belg* 30, 579-94 (1930). WALLACE R. BRON

The preparation of ellagic acid. FRITZ ZETSCHE and MARGUERITE GRAEF. *Helv. Chim. Acta* 14, 240-2 (1931) — Ellagic acid (I) was obtained in 4% yield by acidifying the hydrolysis product of Turkish tannin with an equal wt. of Na_2SO_3 crystals. It forms an addn. compd. with 2 mols. of pyridine. Acetylation of I or its pyridine compd. with Ac_2O yields tetraacetyl-ellagic acid, m. 317-0°. The hydrolytic products also contained 3.3% of 1,2,3 $C_6H_3(OH)_3$. The residue from the aq. ext. of the tannin contained 24% of cellulose. Other tannins also yield I. L. M. LEVINE

α -Bromoisovaleric acid. C. S. MARVEL AND V. DE VIGNEAUX. *Org. Syntheses* XI, 20-2 (1931) — $Me_2CHCH_2(CO_2Et)_2$ is sapon., the free acid liberated by HCl, brominated and then heated at 125-3° until evolution of CO_2 ceases, the yield of $Me_2CHCHBrCO_2H$ is 55-66%. C. J. WEST

Action of ethylmagnesium halides on ethyl diethylcyanoacetate. A. MAYER. *Compt. rend.* 192, 363-5 (1931) — $EtMgBr$ or $EtMgI$ reacts with $NCCEt_2CO_2Et$ (I) either at the CN or CO_2Et group but not with both groups in the same mol. The mixt. of products contains Et_2CO , $Et_2CHCOCEt_2CO_2Et$ and Et_2CHCO_2Et , these compds. were also found when $PhMgBr$ was used. Besides these compds. 3-cyano-4-hydroxy-3,4-diethylhexane (II), bp 132-3°, Et_2CHCN and Et_2COH were also produced. The last was isolated as the *allophanate* (reaction product with $HClNO$), m. 152-3°. I was transformed into 3,3,4-triethyl-4-cyanohexane, bp 105°, by reaction with $SOCl_2$. The work confirms the hypothesis that the ester group in I undergoes enolization under the influence of the Grignard reagent, furthermore, the great complexity of the reaction is considered by M. to be due to the inertia and fragility imparted to the mol. by the tetra-substituted C atom. I. M. LEVINE

Action of ethylmagnesium bromide on *N*-diethylchloroacetamide. SON PHOU TI. *Compt. rend.* 191, 943-5 (1930) — The principal product of the action of 3 mols. of $EtMgBr$ on 1 mol. of $ClCH_2CONEt_2$ was neither a chloroamine nor a chloroketone as might have been expected but 2-diethylamino-2-ethyl-1-ethanol, $Et_2C(NEt_2)CH_2OH$ (I) (20% yield), bp 90°, picrate, m. 94°, chloroplatinate, m. 140-1°. Heating I for 3 hrs. on a steam bath with excess Ac_2O gave its acetate, bp 97-8°, picrate, m. 130-1°. Other products of the reaction were considerable Et_3NH , about 1% of an ethylene base, C_6H_5N , with a picrate, m. 74° and a chloroplatinate m. 142-3°, $Et_3NCH_2CONEt_2$, bp 125-6°, with a picrate, m. 121° (Hahn and Loos report 133°, C. A. 13, 725) and a chloroplatinate, m. 164-5°, and the major product was a syrupy, yellowish, basic oil, $C_6H_5N_2O_2$, bp 160-2°, which did not react with Ac_2O , was not sapon. by 45% HBr, gave no cryst. derivs. and could not be identified. C. H. FEE

Chemistry of thiocholine halides (trimethylthioethylammonium halides). II. Thiocholine chloride and its derivatives. TAICHI HARADA *Bull. Chem. Soc. Japan* 6, 25-8 (1931), cf *C. A.* 24, 2428—Forty g of $(\text{CH}_3\text{Cl})_3$ was condensed with NMe_3 , the chlorocholine chloride (I), $\text{Me}_3\text{NCICH}_2\text{CH}_2\text{Cl}$, formed being washed with Et_2O and dried over CaCl_2 . The 40 g of crude product, recrystd from 95% alc melted about 242° , with decompn above 163° . Six g of I was heated with 5 g of 2-thiouracil together with 25 cc. of H_2O in a bomb tube at 150° for 2 hrs. After filtration from the uracil, the filtrate was neutralized with NH_4OH and on concn more uracil was removed. Fractional crystn from a cold mixt of BzOH and EtOH gave 0.5 g thiocholine chloride (II), $\text{Me}_3\text{NCICH}_2\text{CH}_2\text{SH}$, m 238° with decompn above 200° . II reacts with AgCl to form $\text{Me}_3\text{NCICH}_2\text{CH}_2\text{SAG AgCl}$ sol in H_2O to the extent of 1.215 g per 1 at 37° . It decompn above 70° , m 228° . C. R. ADDINALL.

α -Aminoisobutyric acid. H. T. CLARKE AND H. J. BEAN *Org. Syntheses* XI, 4-6 (1931).—Details are given for the prepn of $\text{Me}_2\text{C}(\text{NH}_2)\text{CO}_2\text{H}$ from Me_2CO , NaCN and NH_4Cl , giving $\text{Me}_2\text{C}(\text{OH})\text{CN}$, with NH_3 this yields $\text{Me}_2\text{C}(\text{NH}_2)\text{CN}$, hydrolyzed by HBr to the acid HBr salt, from which the free acid is obtained in 30-3% yield by $\text{C}_2\text{H}_5\text{N}$. C. J. WEST.

The action of formaldehyde on amino acids with special reference to the formation of amines. LAWRENCE ZELENY AND ROSS A. GORTNER *J. Biol. Chem.* 90, 427-41 (1931).—On boiling casein, alanine, cystine, glutamic acid or tyrosine with HCHO in 20% HCl , 12-40% of the total N is converted first into primary, then into secondary and tertiary, amines, which were estd by Weber and Wilson's method (*C. A.* 12, 1977). NH_4Cl behaves similarly. The end product, which was reached more rapidly by autoclaving with HCHO at 180° in 5% HOAc , was in all cases NMe_3 , identified by the compd with Reinecke's salt or by the chloroplatinate. The action of HCHO is therefore one of deamination followed by methylation as suggested by Werner (*C. A.* 11, 3243), but occurring through some intermediate non-volatile, non-amino compd. This compd, which may be a CH_2 deriv, was not isolated, but after the amino-acid N has reached a const. value, the amt. of volatile N continues to increase on boiling. K. V. TIMMANN.

N-Alkylation of β -amino esters. J. DÉCOMBE *Compt. rend.* 191, 945-7 (1930).—Alkylation of β -amino esters by treatment with alkyl iodides proceeds unsatisfactorily because it is impossible to fractionate the resulting mixt by distn and because there is decompn. of the parent amine through spitting out of NH_3 with resultant formation of an unsatd ester. However, 70% yields of alkylamino esters may be obtained by the catalytic hydrogenation in EtOH of equimol. mixts of the β -amino ester and the appropriate aldehyde, with PtO_2 (Adams) as a catalyst. The reaction doubtless proceeds through a condensation of the amine with the aldehyde and reduction is accompanied by the formation of a certain amt of resin. The following compds were prepd by condensing the proper β -amino ester with AcH : $\text{MeCH}(\text{NH}_2\text{Et})\text{CH}_2\text{CO}_2\text{Et}$ (I), b. $74-5^\circ$ (HCl salt, m 125°); $\text{MeCH}(\text{NH}_2\text{Et})\text{CH}_2\text{CO}_2\text{H}$, m $169-70^\circ$; $\text{PrCH}(\text{NH}_2\text{Et})\text{CH}_2\text{CO}_2\text{Et}$, b. $110-1^\circ$ (HCl salt, m 79°); $\text{C}_4\text{H}_9\text{CH}(\text{NH}_2\text{Et})\text{CH}_2\text{CO}_2\text{Et}$, b. $130-2^\circ$ (HCl salt, hygroscopic, m about 60°); $\text{C}_4\text{H}_9\text{CH}(\text{NH}_2\text{Et})\text{CH}_2\text{CO}_2\text{H}$, m. $140-1^\circ$. The method is not so satisfactory for the introduction of 2 alkyl groups. With aldehydes of low mol wt, a mixt. of amino esters is obtained which cannot be fractionated. Aldehydes of higher mol wt yield a small quantity of the dialkyl deriv. $\text{MeCH}(\text{NEtBu})\text{CH}_2\text{CO}_2\text{Et}$, from PrCHO and I, b. $118-9^\circ$ and forms a viscous HCl salt but a microcryst. chloroplatinate. Prepn. of dialkyl derivs by treatment of the monoalkyl compds. with RI is also unsatisfactory (The b. p. at 17 mm. of $\text{MeCH}(\text{NH}_2)\text{CH}_2\text{CO}_2\text{Et}$, incorrectly given in *C. A.* 24, 2110, is $68-9^\circ$). C. II. PEST.

N-Acyl derivatives of alanine. Resolution of externally compensated *m*-nitrobenzoylalanine. WM. M. COLLES AND CHARLES S. GIBSON. *J. Chem. Soc.* 1931, 279-85.—The *N*-acyl derivs of alanine (I) were studied not only as a means of isolating and identifying it but also because the nitroacyl derivs. may serve as convenient starting materials for prep. arsonic acids for the study of the relative chemotherapeutic activities of *dl*- and active isomers. *dl*-*m*-Nitrobenzoyl deriv (II) of I, m. $163-4^\circ$; at 20° 0.38 g dissolve in 100 cc. H_2O , 0.37 g in 100 cc. Et_2O and 0.43 g in 100 cc. CHCl_3 ; Ag salt, needles; Et ester, m 89° ; Me ester, m $110-1^\circ$; NH_4OH gives the amide, m. $189-90^\circ$. Reduction of II gives the NH_2 deriv, whose HCl salt m. $150-2^\circ$ (decompn.). *dl*-*o*-Nitrobenzoyl deriv (III) of I, m. $165-6^\circ$, slightly less sol in H_2O and Et_2O than II; *o*- NH_2 deriv, does not m 270° . *dl*-*p*-Toluy deriv (IV), m. $188-9^\circ$ (95% yield). *dl*-*m*-Nitrobenzenesulfonylalanine (V), from I in N KOH and $\text{m O}_2\text{NC}_6\text{H}_4\text{SO}_2\text{Cl}$ in C_6H_6 , added at regular intervals, yellow, m. $158.5-9^\circ$ (89% yield). *dl*-4-Nitrotoluene-2-sulfonylalanine (VI), crystg. with 1 mol. H_2O , m. $91-7^\circ$; anhyd., it m. $125.5-6.5^\circ$. *dl* Cin-

namoylalanine (VII), m. 19°-7° (30% yield), in alk. soln. it rapidly decolorizes KMnO_4 , giving Bril 1 (15 g.) in 32.4 cc. 0.9% N NaOH and 1000 cc. H_2O , treated with 11.92 g. quinine and then with 10 cc. EtOH, gives 17 g. of the quinine salt of I-I, converted to a glass-like mass at 125° crystals with 2 mols. H_2O , $[\alpha]_{\text{D}}^{25} -137.1^\circ$; the mother liquor was freed of alal and then treated with brucine, giving 58.2 g. of the brucine salt of d-I, cryt'g with 3.5 mols. H_2O $[\alpha]_{\text{D}}^{25} -9.1^\circ$ (EtOH, c 0.5511) I-L, m. 158°, $[\alpha]_{\text{D}}^{25} -44.7^\circ$ (NH_4 salt in H_2O c 1.342) 5.87° (EtOH, c 1.351). Et ester, m. 104-5°, $[\alpha]_{\text{D}}^{25} 6.91^\circ$ (EtOH, c 1.455), in EtOH KOH (0.0832 g. KOH in 50 cc. EtOH) α_{D}^{25} changes from 0.17 to -1.05° in 26 hrs. d-I, m. 158°, $[\alpha]_{\text{D}}^{25} 44.7^\circ$ (NH_4 salt in H_2O , c 1.355), -5.62° (EtOH c 1.355). Me ester, m. 125°, $[\alpha]_{\text{D}}^{25} -12.7^\circ$ (EtOH, c 1.429); amide, m. 193-4°, $[\alpha]_{\text{D}}^{25} 21.2^\circ$ (EtOH, c 0.235). The relative rates of hydrolysis of the di-N-acyl derivs. were det'd. by gently heating 3 hrs. with const.-boiling HCl, neutralizing the hydrolysis mixt. to pH 8 and titrating the liberated I by Sørensen's method to pH 10. The % hydrolysis is given: Br, 96, IV, 95, VII, 92, phthalyl, 76, III, 47, II, 91; $\text{p-O}_2\text{NC}_6\text{H}_4\text{CO}$ deriv., 9°; PhSO_2 deriv., 14, $\alpha\text{-C}_6\text{H}_4\text{SO}_2$ deriv., 65, β isomer, 18 (45% after 9.5 hrs.); $\text{p-MeC}_6\text{H}_4\text{SO}_2$ deriv., 17, V, 21, VI, 21.4 hrs.

C. J. WEST

d-Glyceraldehyde E. J. WITTEMACY, WM. LLOYD EVANS, HENRY HASS AND E. F. SCHROEDER. *Org. Syntheses* XL, 50-1 (1931)—Hydrolysis of 50 g. $\text{HOCH}_2\text{CH}(\text{OH})\text{CH}(\text{OEt})_2$ with 500 cc. 0.1 N H_2SO_4 at 20° for about 1 week gives 50% of $\text{HOCH}_2\text{CH}(\text{OH})\text{CHO}$, m. 137-0°.

C. J. WEST

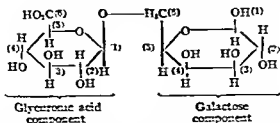
d-Glyceraldehyde acetal. E. J. WITTEMACY, WM. LLOYD EVANS, HENRY HASS AND E. F. SCHROEDER. *Org. Syntheses* XL, 52-3 (1931)— $\text{CH}_3\text{CH}(\text{OH})\text{CH}(\text{OEt})_2$ and KMnO_4 in H_2O at 5° give 67% of $\text{HOCH}_2\text{CH}(\text{OH})\text{CH}(\text{OEt})_2$, b. 120-1°. C. J. W.

The two forms, cis and trans, of isobutyldiacetone and the isomeric unsaturated β,γ -ketone R. HENLMANN. *Bull. soc. chim.* [4], 49, 75-80 (1931)—Eccott and Linstead have shown (C. A. 24, 3701) that the condensation of Me_2CHCHO with Me_2CO gives the trans-modification of $\text{Me}_2\text{CHCH}=\text{CHCOMe}$ (I) which on boiling with 20% H_2SO_4 is isomerized into $\text{Me}_2\text{C}=\text{CHCH}_2\text{COMe}$ (II). In prep'g $\text{Me}_2\text{CHCH}=\text{CHNHN} \cdot \text{CMe}_2\text{CH}_3$ (III) by the action of $(\text{NH}_4)_2\text{H}_2\text{O}$ on I, H. obtained a fraction which on hydrolysis with 20% H_2SO_4 gave II. Although II might result from the hydrolysis of the amine produced by the spontaneous oxidation of III (C. A. 24, 851), thus regenerating the 2 stereoisomeric forms of I, and by the consequent isomerization of the trans-modification, there does not seem to be absolute identity between the residues from the prep'n. of III and the products obtained by controlled oxidation of the pyrazoline. In the 1st case the oil obtained consists almost entirely of III whereas in the 2nd case the principal product is I. H. is still of the opinion that I and II are formed simultaneously in the condensation of Me_2CHCHO with Me_2CO and proposes to return to this matter at a later date.

C. R. ADDINALL

Compound uronic acids. Structure of the aldobionic acid from gum arabic. SYDNEY WM. CHALLINOR, WALTER N. HAWORTH AND EDMUND L. HIRST. *J. Chem. Soc.* 1931, 208-65—Aldobionic acid from gum arabic, with $[\alpha]_{\text{D}}^{25} 2.0^\circ$ (Ca salt), on methylation with MeSO_3 and 30% NaOH and finally with Ag_2O and MeI , gives a mixt. of α - and β -forms of Me heptamethylaldonate (Me ester of hexamethyl-6-glycyronido- β -methylgalactonate), b.p. 155°, $n_{\text{D}}^{25} 1.4715$, $[\alpha]_{\text{D}}^{25} -3.5^\circ$ (H_2O , c 1.55), -24° (CHCl_3 , c 2.0), which is not redcing. after 14 days at 50-60° about 20% of the mixt. seps. as the cryst. β -deriv., m. 85°, $[\alpha]_{\text{D}}^{25} -21^\circ$ (H_2O , c 0.67), -43° (CHCl_3 , c 0.55). Hydrolysis of the liquid ester was effected by heating 3.5 hrs. at 100° in 7% HCl, $[\alpha]_{\text{D}}^{25}$ changing from -1° to 63.5° in that time. The 2 hydrolysis products were isolated in 87-8% yields: 2,3,4-trimethylgalactose, pale yellow syrup, $n_{\text{D}}^{25} 1.4727$, $[\alpha]_{\text{D}}^{25} 83^\circ$ (H_2O , c 0.6); further methylation gives 2,3,4,6-tetramethyl- β -methylgalactonate. Br- H_2O at 25° for 24 hrs. gives 2,3,4-trimethyl- β -galactonolactone (after heating 2 hrs. at 100° and 12 mm.), which was characterized as the phenylhydraide, m. 163-7°; the rate of hydrolysis of the lactone is given, becoming cryst. in 8 hrs. Oxidation of the lactone gives 2,3,4-trimethylmalic acid, whose di-Me ester m. 98°, $[\alpha]_{\text{D}}^{25} 35^\circ$ (H_2O , c 0.6). The other hydrolytic product is 2,3,4-trimethylglyceronic acid, pale yellow, viscid syrup, $n_{\text{D}}^{25} 1.4709$, $[\alpha]_{\text{D}}^{25} 55^\circ$ (equil. value in H_2O , c 0.63) (this value is for a product that has not been dried); further methylation gives a mixt. of the α - and β -forms of 2,3,4-trimethylmethylglyconate, the β -form m. 133°, $[\alpha]_{\text{D}}^{25} -28^\circ$ (H_2O , c 0.4), $[\alpha]_{\text{D}}^{25} -63^\circ$ (CHCl_3 , c 0.6). This was also obtained from glyceronic acid with MeSO_3 and alkali. In N HCl at 100° the velocity of hydrolysis of the glyconate was almost identical with that of β -Me glucoside under

similar conditions. Thus, the structure of the aldobronic acid from gum arabic appears to be.



It is not certain whether α - or β -glucuronic acid is involved in the biase linking, although preliminary evidence favors the β -linking.

Ethyl ethylenetetracarboxylate. B. B. CORSON AND W. L. BENSON. *Org. Syntheses* XI, 36-8(1931)— $\text{BrCH}(\text{CO}_2\text{Et})_2$ (123 mols.) and 19 mols. anhyd. Na_2CO_3 , heated 3 hrs. at $150-160^\circ$, give 75-80% of $\text{C}(\text{CO}_2\text{Et})_4$, b_p 197° , b_m 203° , b_n 210° , b_d 221° , b_{15} 234° , m 52.5-53.5°. The yield in 95% EtOH is reported for 0 to 35.5%.

Condensations giving diethyl dicyanoglutaconate. YOSHITAKI UREKIMURA. *Bull. Chem. Soc. Japan* 6, 29-31(1931)—The compd. termed di-Et dicyanoglutaconate, $\text{C}_{10}\text{H}_{10}\text{O}_4\text{N}_4$ (I) (*C. A.* 21, 3889), is now considered to be the semihydrate ($\text{C}_{10}\text{H}_{10}\text{O}_4\text{N}_4 \cdot \text{H}_2\text{O}$) (II). The acidification of di-Et sodio dicyanoglutaconate gives II and so attempts were made to form I by direct condensation rather than by proceeding through this intermediate Na compd. Two such condensation processes were carried out in which the formation of I was proved by the pptn. of II on acidifying the alkali ext. of the crude products, e. g., the condensation of $\text{HOCH}_2\text{C}(\text{CN})\text{CO}_2\text{Et}$ with $\text{CNCH}_2\text{CO}_2\text{Et}$ by the action of Ag_2O and the condensation of $\text{ClCH}_2\text{C}(\text{CN})\text{CO}_2\text{Et}$ with $\text{ClCH}(\text{CN})\text{CO}_2\text{Et}$ by the use of metallic Cu.

Mesaconic acid. R. L. SHRINER, S. G. FORD AND L. J. ROLL. *Org. Syntheses* XI, 74-5(1931)—Citraconic anhydride (100 g.), 100 cc. H_2O and 150 cc. dil. HNO_3 evaporated to red fumes, give 43-52% of mesaconic acid, m $273-5^\circ$.

Ethyl fumelate. ADOLF MÜLLER AND ERICH RÖHL. *Org. Syntheses* XI, 42-5(1931)—Details are given of the reduction of $\alpha\text{-HOCH}_2\text{CH}_2\text{CO}_2\text{H}$ with iso-AmOH and Na and the esterification of the acid; yield of ester, b_p $153-6^\circ$, $35-57^\circ$.

Citraconic anhydride and citraconic acid. R. L. SHRINER, S. G. FORD AND L. J. ROLL. *Org. Syntheses* XI, 28-9(1931)—Itaconic anhydride, d. rd. rapidly at atm. pressure, gives 68-72% of citraconic anhydride, soln. in slightly more than the necessary amt. of H_2O gives 94% of the acid.

Fumaric acid. NICHOLAS A. MINAS. *Org. Syntheses* XI, 46-8(1931)—Oxidation of furalural with NaClO_2 (V_2O_5 as catalyst) gives 50-55% of fumaric acid.

Cyanogen bromide. W. W. HARTMAN AND E. E. DRAGER. *Org. Syntheses* XI, 30-1(1931)—Details are given of the reactions of NaCN and Br , giving 73-83% of BrCN .

Isopropyl thiocyanate. R. L. SHRINER. *Org. Syntheses* XI, 92-3(1931)—Directions are given for the prepn. of Me_2CHSCN from iso-PrBr and NaSCN in 76-97% yields.

Oxidation. IV. Action of ferric chloride and hydrogen peroxide on 5-alkylthiosemicarbazones. Formation of triazoles. SATISH CHANDRA DE AND TARANI KANTA CHAKRAVORTY. *J. Indian Chem. Soc.* 7, 875-8(1930), cf. *C. A.* 24, 4781.—By oxidation of 5-alkylthiosemicarbazones of aldehydes triazole formation is excluded and triazoles formed. Yields with FeCl_3 were poor but with H_2O_2 the reaction proceeds smoothly. 5-Phenyl-3-methylthio-1,2,4-triazole, m 164° , was obtained by condensing $\text{NH}_2\text{C}(\text{SMe})\text{NNH}_2$ and BzH in alc. followed by H_2O_2 oxidation without isolation of the intermediate compd. 5-Phenyl-3-ethylthio-1,2,4-triazole, m 166° . Benzylthio-5-methyl-4-phenylthiosemicarbazone, m $66-7^\circ$; HI soln. m 153° . 4,5-Diphenyl-3-methylthio-1,2,4-triazole, m $155-6^\circ$. Benzylthio-5-ethyl-4-phenylthiosemicarbazone, m 78° ; HI soln. m 118° ; HCl soln. m 178° (decomp.). 4,5-Diphenyl-3-ethylthio-1,2,4-triazole, m 145° . Benzylthio-5-methyl-4-ethylthiosemicarbazone, m 62° ; HI soln. m 159° (decomp.); HCl soln. m 194° (decomp.). 4-0-Tolyl-5-phenyl-3-methylthio-1,2,4-triazole, m 130° . Benzylthio-5-ethyl-4-ethylthiosemicarbazone, HI soln. m 123° ; HCl soln. m 157° . 4-0-Tolyl-5-phenyl-3-ethylthio-1,2,4-triazole, m 107° . Benzylthio-5-methyl-4-ethylthiosemicarbazone, m 71° ; HI soln. m 180° (decomp.); HCl soln.

m. 130° (decompn). 4-*p* Toly 5-phenyl-3-methylisoxol-1,2,4 triazole, m. 176°. Benzylidene 5-ethyl-4-*p*-tolylthiosemicarbazone, III salt, m. 165°; HCl salt, m. 158°. 4-*p*-Tolyl-5-phenyl-3-ethylisoxol-1,2,4 triazole, m. 148° IOSTRA DFE SNELL

Carbamidosulfonic acids. PAUL BAUMGARTEN AND ILSE MARGGRAFF. Ber. 64B, 391-9 (1931).—N-Pyridiniumsulfonic acid and urea give $\text{NH}_2\text{CONHSO}_3(\text{C}_5\text{H}_4\text{N})$ (I) or $\text{CO}(\text{NH}_2\text{SO}_3)_2(\text{C}_5\text{H}_4\text{N})_2$ (II) depending on the mol proportions. The free acids are unstable. Both form stable metal salts. The salts of I in acid soln decomp into urea and H_2SO_4 in neutral or acid soln (i.e., the mono- or divalts) decomp into NH_3 , CO_2 and $\text{H}_2\text{NSO}_3\text{Na}$. The salts of II in acid or neutral soln decomp into urea and H_2SO_4 and also into CO_2 and $\text{NH}_4\text{SO}_3\text{H}$. The neutral divalts decomp in dil soln into CO_2 and $\text{H}_2\text{NSO}_3\text{Na}$ and in solns over 2% into urea and NaHSO_4 . The tetravalts are very stable. These reactions are analogous to those of nitrourea V. F. HARRINGTON

Preparation of xanthates and other organic thiocarbonates. LAWRENCE S. FOSTER. Dept. Mining and Met. Research, Univ. of Utah, Tech. Paper 2, 8 pp. (1928).—The history of the xanthates is reviewed and methods are described for the prepn. of iso-Am, hexyl, heptyl, octyl, nonyl, lauryl and other xanthates, and K alkyl monothiocarbonate, Na Bu and K iso-Am trithiocarbonates ALDEN H. LIMEY

A catalyst for the autoxidation of uric acid. M. PRÉREJACQUE. Compt. rend., 191, 949-51 (1930).—Autoxidation of uric acid (I) in KOH can be reduced from 15 hrs. to 10 min. by the addn. of 5% of active C, thus eliminating side reactions. Active C also catalyzes the alk. oxidation of 1-methyl-, 1,3-dimethyl-, hydroxymethylene- and 7-methyl- (II) uric acids although total oxidation of II in the absence of the catalyst appears impracticable. Autoxidation of allantoin (III) by active C in alk. soln is slow and gives 30% of K oxonate (IV). In the absence of the catalyst, this step cannot be observed on account of the preponderant hydrolysis of III to allantoinic acid. There is no autoxidation of the other purines (xanthine (V), hypoxanthine, caffeine, theobromine, etc.) even in the presence of C, hence, pure V is easily obtained by agitating the crude product from the reduction of I with O and C. Autoxidation of I takes between 1 and 2 atoms of O per mol of I—most at low temps., least at high. Acidification after oxidation at ordinary temps. gives a mixt. of III and IV, the amt. of IV increasing as the amt. of O absorbed increases. When I is oxidized at 0° in the presence of 1% of C, there is no evolution of NH_3 and acidification after filtering out the C ppt. oxonamide (VI). Perhaps the oxidation proceeds as follows: a peroxide of I (RO_2) is formed, this can decomp into RO and O (C catalyzes this decompn.), the reagents present decomp the RO_2 and RO into III and VI. C. H. FREY

Acetone compounds of the sugars and their derivatives. XVII. Conversion of monoacetoneglucose into a new amino- and anhydroglucose. Also a contribution to acyl migration. HERZ OHLÉ AND RUDOLF LICHTENSTERN. Ber. 63B, 2905-12 (1930); cf. C. A. 24, 1351.—In connection with the tautomerism problem of the sugars it was of interest to prep the still unknown glucosyl 5-amine. 6-*p*-Toluenesulfonyl-6-benzoyl-acetoneglucofuranose seemed to be a suitable starting material for this purpose, and after all attempts partially to hydrolyze it to the toluenesulfonylacetoneglucofuranose (I) had failed (the $\text{MeC}_6\text{H}_4\text{SO}_2$ group was split off along with the Bz group), it was treated directly with NH_3 in MeOH. The reaction might conceivably proceed in various directions. Presumably, the Bz group is first split off and the intermediate I might then (1) yield acetoneglucosyl 5-amine (II) by simple exchange of the $\text{MeC}_6\text{H}_4\text{SO}_2$ group for NH_2 , or (2) lose $\text{MeC}_6\text{H}_4\text{SO}_2\text{H}$ to form 5,6-anhydroacetoneglucose (III) and add NH_3 to give acetoneglucosyl-6-amine (IV). Moreover, a Walden inversion at the 5-C atom might occur in either (1) or (2), giving (3) acetoneidrosyl 5-amine (V) or (4) acetoneidrosyl-6-amine (VI) through the corresponding ethylene oxide. As a matter of fact, the reaction is more complicated and has not yet been entirely cleared up but the results thus far obtained are published now because of the appearance of Josephson's paper (C. A. 24, 5285) in which he announces that he is going to attempt to prep acetoneglucose 5-phosphoric acid. The reaction with NH_3 in MeOH proceeds only slowly at room temp. and the rotation of the mixt. does not become const. even after 3 weeks. If the reaction is interrupted after 14 days, there can be isolated the *p*-toluenesulfonate (VII) of an acetonehexoamine (VIII) and a syrupy, N-free, neutral substance (IX). VII is different from the toluenesulfonate of V. HNO_3 gives an acetoneanhydrohexose (X) different from either III or the 3,6-compd., nor, from its mol wt. and low m. p., can it be a dimer of either of these compds. It does not add NH_3 , even at 100°. Nor is the ring opened by acids, not even by $\text{N H}_2\text{SO}_4$ at 100°; there is merely obtained the free anhydrosugar (XI) which, like 3,6-anhydroglucose, is sweet and shows no mutarotation. This stability toward hot dil. mineral acids makes it very improbable that the O bridge is between C atoms 3 and 5. It is therefore assumed that XI is 3,6-anhydrohexose and

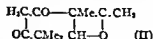
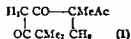
VIII is *acetoneidoxyl-6-amine*. Of the 4 possible reactions mentioned above, apparently only (4) takes place. Some other reaction also takes place at the same time, however, as indicated by the formation of IX which on acetylation gives an *acetyl-*p*-toluenesulfonylacetonehexose* (XII) different from the 3,5,6-compd and characterized by a surprising resistance to dil mineral acids. Whatever the structure and configuration of IX may be, its formation is an entirely new phenomenon in the chemistry of the sugars, since the starting material contained only one $\text{MeC}_6\text{H}_4\text{SO}_2$ group, one part of it must have been toluenesulfonated under the influence of the $\text{NH}_2\text{-MeOH}$, at the expense of another part; the 5-position of the $\text{MeC}_6\text{H}_4\text{SO}_2$ group is thought to stand in some causal relationship with this reaction. VII (15 g from 44 g of the toluenesulfonylbenzoylacetoneglucose), m 173–4° (decompn.), $[\alpha]_D^{25} -23.49^\circ$ (H_2O , c 3.150) (the salt of V decomps. 176–7°, $[\alpha]_D^{25} -7.02^\circ$ (H_2O , c 5.01)). XII, m 112°, does not reduce Fehling soln until after treatment with concd H_2SO_4 or heating with $\text{NH}_2\text{-MeOH}$ at 100°, $[\alpha]_D^{25} 2.87^\circ$ (CHCl_3 , c 2.436) (the 3,5,6-compd. m 92°, $[\alpha]_D^{25} -23.76^\circ$ (CHCl_3 , c 4.347)). H_2SO_4 (0.05 N) completely splits off the Me_2CO from VIII only after 8 days at 37°, $[\alpha]_D$ in a 1-dm tube (c 10) falling from -1.0° to -0.25° . Direct treatment of the hydrolyzate (after distg off the Me_2CO split off and neutralizing the H_2SO_4 with KOAc) gives with PhNHNH_2 , *phenylhydrazine-p-toluenesulfonate*, $\text{C}_{11}\text{H}_{11}\text{O}_5\text{SN}$, H_2O , m 183–4° (decompn), identical with Ohle and v. Vargha's supposed toluenesulfonate of the phenylhydrazone of glucosyl-6-amine. *p-Nitrophenylhydrazone* of XI, dark red brown, m around 227°. *Tri-Bz deriv* of VIII, m 192°, $[\alpha]_D^{25} -19.30^\circ$ (CHCl_3 , c 2.280), one of the Bz groups is held very loosely and the compd is partially alcoholized when crystd from alc. *Tri-Ac deriv* of IV, m 86°, $[\alpha]_D^{25} 18.35^\circ$ (CHCl_3 , c 2.452), *tri-Bz deriv*, m 198–9°, $[\alpha]_D^{25} -76.43^\circ$ (CHCl_3 , c 2.682), is not appreciably alcoholized, *di-p-toluenesulfonyl deriv*, m 172° (decompn), $[\alpha]_D^{25} -2.54^\circ$ (CHCl_3 , c 3.145). *Hexabenzoylglucosyl-6-amine*, needles with 1 mol PhMe , reduces Fehling soln, $[\alpha]_D^{25} 22.89^\circ$ (CHCl_3 , c 0.830). *Acetone-3,6-anhydro-dextrose* (X) (yield, 80%), m 105°, $b.p.$ 147–8° (bath temp), $[\alpha]_D^{25} 24.94^\circ$ (H_2O , c 2.566), $\text{mol. wt. in H}_2\text{O}$ 202.1 (c 1.022) (3,6-anhydroacetoneglucose, m 56–7°, $[\alpha]_D^{25} 29.33^\circ$ (H_2O)); III, m 133.5°, $[\alpha]_D^{25} -26.5^\circ$ (H_2O). XI, m 105–6°, is markedly hygroscopic, strongly reduces hot Fehling soln, $[\alpha]_D^{25} 25.36^\circ$ (H_2O , c 3.510) (3,6-anhydroglucose, m 118°, $[\alpha]_D^{25} 57.39^\circ$ (H_2O , c 2.883)).

C. A. R.

Report (No. 93) of the most important work in the field of pure sugar chemistry published during the first half of the year 1930. E. O. v. LIPPMAN, *Deut. Zuckerind.* 55, 1104, 1131–3, 1178–9, 1204 (1930).—A comprehensive review.

J. L. LEETE

Essential oil of *Backhousia angustifolia*. III. Constitution of angustione and dehydroangustione. ROBERT S. CAHN, CHARLES S. GIBSON, ARTHUR R. PENFOLD AND JOHN L. SIMONSEN, *J. Chem. Soc.* 1931, 286–94, cf. *C. A.* 24, 3998.—The constitution ascribed to angustione (I) is further confirmed by the action of HBr , which gives 1,1,3-trimethylcyclohexane-4,6-dione, m 130–1°, and by its oxidation with KOBBr to $\text{HO}_2\text{CCHMeCH}_2\text{CMe}_2\text{CO}_2\text{H}$, m 91–2°, $[\alpha]_D^{25} -12^\circ$ (H_2O , c 1.546), oxidation with KMnO_4 gives a ketonic acid, further oxidized to the above acid with KOBBr . Dehydroangustione (II), $p\text{-MeC}_6\text{H}_4\text{NH}_2$ and ZnCl_2 , heated 2 hrs at 140–50°, give the *p-toluidino deriv*, m 63–5°. Heating II with HBr on the H_2O bath for 4 hrs gives 1,1,3-trimethyl- Δ^4 -cyclohexene-4,6-dione (III), m 156–7°. Reduction of II gives *di*-angustione, $b.p.$ 127°, $d_4^{20} 1.083$, $n_D^{20} 1.5087$ (identical with the natural ketone, except for the value of n , the previous sample was contaminated with II), the amino deriv, m 139–40°, is formed by adding NH_4OH to the EtOH soln.; the *s* (or *c*)-oxime, m 57–8°. Oxidation of II with NaOBr gives a mixt. of a compd. m 83.5°, identified as the anhydride of *cis*- α,β,γ -



trimethylglutaconic acid (IV), m 125° (on bromination, the product evolves HBr , giving the lactone of β -bromo- γ -hydroxy- α,α,γ -trimethylglutaric acid, m 147–8°) and *trans*- β -hydroxy- α,α,γ -trimethylglutaric acid, m 156–7°. Oxidation of III with NaOBr gives as the main product IV; there also results a small quantity of a compd., $\text{C}_{11}\text{H}_{19}\text{O}_4$, m 85–6°, which gives a purple color with FeCl_3 , this is unstable and passes into a red oil on standing, it may be $\text{OCCHMeCH}_2\text{CMe}_2\text{CO}$ or $\text{HOC}(\text{OH})\text{CMe}_2\text{CHCMe}_2$.

Note on the parachors of angustione and dehydroangustione. OWEN J. EVANS AND

FREDERICK G. SOPER *Ibid* 289-90—*d*₄ 1.51-2°, γ_{20} 34.9, P 442 (calcd., 462.1), Π gave γ_{20} 36.7, P 435 (calcd., 439.5), the d_{20}^{25} were 1.083 and 1.103, resp. The temp. coeff. of the surface tension for I is -0.072 dyne/cm. per degree, d_{20}^{25} 1.074, Ramsay-Shields const. 2.05, this indicates that I is a normal liquid, so that the neg. anomaly cannot be attributed to assocn. The reason for this is discussed C. J. WEST

Model experiments for the theory of alcoholic fermentation. III. Degradation of α -diacetonelfructosulfuric acid. HEINZ OHLE AND GEORG COUTSICOS *Ber* 63B, 2912-27(1930), cf. C. A. 24, 1350—It was shown in Paper I that β -diacetonelfructose-1-sulfuric acid (I) with anq. neutral KMnO_4 gives the tri K salt of 2,3-acetone- β -1-furtondiacid-1-sulfuric acid (II) (the mother substance, $\text{HOCH}_2\text{C}(\text{OH})(\text{CH}(\text{CO}_2\text{H}))_2$), being designated as furtondiacid) The prepn. of this salt has since been materially improved and simplified and it is now quite readily available. A weak point in the proof of its constitution was the detection and quant. estn. of the $\text{HOCH}_2\text{CO}_2\text{H}$ formed in its hydrolysis but this gap has now been filled in. On the basis of the fact that $\text{HOCH}_2\text{CO}_2\text{H}$ forms a difficultly sol. basic Pb salt a quant. method for detg. this acid, sufficiently accurate for the present purpose, was devised and it was found that the above tri K salt does in fact give 1 mol. $\text{HOCH}_2\text{CO}_2\text{H}$, as previously assumed. That the yield of MeCOCHO does not exceed 80% is probably due to a small part of it undergoing a Cannizzaro reaction, with formation of AcCO_2H and AcCH_2OH ; the latter, though not isolated, was shown to be present by its characteristic odor. The reaction was next applied to α -diacetonelfructose-3-sulfuric acid (III). Whereas with I the formation of II and the point where it is present in greatest concn. can be readily recognized by following the oxidation polarimetrically to the max. in the rotation curve, the curve for III shows no such max. With 12 atoms O the III is completely broken down into optically inactive products, as is demanded by theory if the Me_2CO split off is not attacked. If each mol. of III were at once broken down into the final products (Me_2CO , H_2O , CO_2 , H_2SO_4), before the next mol. underwent the same change, the rotation curve should be a straight line, but, as was to be expected, it is not. To obtain further light on the course of the reaction, the splitting off of the H_2SO_4 was also followed and here again the curve obtained by plotting the no. of mols. H_2SO_4 split off against the no. of atoms of O used was not a straight line. In both cases the max. deviation from a straight line was at a point corresponding to about 6 atoms O, and at this point it was possible to isolate an almost optically inactive product (IV) having the same compn. as II, i.e., the tri K salt of 1,2-acetone- α -1-furtondiacid-3-sulfuric acid, and yielding the same hydrolysis products. Attempts to isolate intermediate products in the formation of II and IV in order to det. at what point of the fructose mol. (C atom 5 or 6) oxidation first begins were unsuccessful, and the study of α -acetonefructose-3-sulfuric acid (V) was therefore undertaken. Since here the 4- and 5-HO groups are free, the V should be more readily attacked by KMnO_4 and the first oxidation phase should use up more KMnO_4 than with III and less KMnO_4 would be left for the further oxidation so that there would be an increased accumulation of the intermediate oxidation product. The max. deviation of the exptl. rotation curve from a straight line appeared at the point corresponding to 5 atoms O, and at this point there were obtained large quantities of IV but also, although in small quantity, the di K salt of 1,2-acetonefructuronic-3-sulfuric acid (VI), showing that, in part, at least, it is the 6-C atom which is first attacked. Whether VI is an intermediate or a by-product in the formation of IV could not be detd. for lack of material. VI, the first deriv. of the hitherto unknown fructuronic acid, the analog in the ketose series of glucuronic acid, is decompd. by boiling dil. HCl into Me_2CO , H_2SO_4 , and CO_2 , but the latter is evolved much more slowly than from II or IV. MeCOCHO and furural were not formed. To det. how the OSO_3H residue influences the course of the oxidation of α - (VII) and β -diacetonelfructose- (VIII) and α -acetonefructose (IX), the oxidation curves of these compds. were also detd. With VIII the rotation curve also changes from $-$ to $+$, with a max. at the point corresponding to about 6 atoms O, so that 2,2-acetone-2,2,3-trihydroxyglutar-4-carboxylic acid is probably formed, but the curve does not pass through α_D^{20} at the point corresponding to 12 atoms O. The same is true of the curve for VII. The only possible explanation for this is that a part of the Me_2CO split off is also oxidized, probably after it has reacted with another cleavage product of the sugar mol. The curve for VII also shows a dispersion of the exptl. values which is not due to exptl. error, as the individual values can be reproduced with great exactitude. It is, therefore, assumed that a whole series of intermediate products is formed through different reactions, one of which, through the furtondiacid, is extraordinarily favored by the introduction of an OSO_3H residue. The curve for IX is parallel to the theoretical (straight line) curve from the points corresponding to 2 up to 9 atoms O, indicating that most of the IX is oxidized directly to Me_2CO , CO_2 , and H_2O without the accumulation in

appreciable amount of any intermediate product. If, however, IX is not a pyroid fructose but a fructofuranose, the curve might be interpreted on the assumption that w 2 atoms O there is an almost quant formation of acetonefructuronic acid, but an ex on a large scale under these conditions yielded only large quantities of unchanged I IX and its 3-sulfate (V) show most strikingly the selective action of the OSO_3H gro which evidently depends on its adjacent position to the masked C O group, the strong electronegative OSO_3H residue stabilizing the hydrate form of the C O group as therefore exerting the same influence as the Cl in $\text{CCl}_4\text{CHO H}_2\text{O}$. By the stabilizat of the hydrate form the union of the Me_2CO is also strengthened and the upper end of t mol is protected from attack by the oxidizing agent. Other electronegative grou must exert the same influence, and, furthermore, oxidation through derivs of the furto diacid type in the glucose series is at most possible only in those glucose derivs hav strongly acid residues on the 2-HO group. It is believed therefore, that a degradati of glucose in the sense of the Neuberg fermentation scheme is possible only by way fructose. K salt of III, from VII and ClSO_3H in $\text{C}_4\text{H}_7\text{N}$, needles with $1/2$ mol. H_2O , decomps about 165° , $[\alpha]_D^{20} -124.3^\circ$ (H_2O , c 1.094), is practically insol in abs. al swells in ordinary alc. and dissolves on heating, the satd cooled soln contg about 0.7 of the salt, more concd solns solidify to a more or less stiff jelly. The alc. in the jellies can be for the most part squeezed out by gentle pressure. The higher alcs beha like EtOH but with MeOH the property of gelatinizing is less marked. Na salt, needl with H_2O , decomps about 170° , $[\alpha]_D^{20} -128.3^\circ$ (H_2O , c 2.144). The salts are stal toward boiling dil alkalis but split off Me_2CO and H_2SO_4 with very dil acids at roc temp. Tri K salt of IV (20 g from 100 g of the K salt of III with 2 N KMnO_4 (6 atom O) on the H_2O bath), prisms with $3\text{H}_2\text{O}$ from MeOH Me_2CO EtOH, $[\alpha]_D^{20} -1.78$ needles with 4.5 H_2O from MeOH, $[\alpha]_D^{20} -2.72^\circ$. K salt of V, obtained by part hydrolysis of the K salt of III with N H_2SO_4 at room temp, decomps about 156° , $[\alpha]_D^{20} -112^\circ$ (H_2O , c 1.847). Na salt, decomps 140° , $[\alpha]_D^{20} -118.6^\circ$ (H_2O , c 2.424). Di-K salt of VI (5 g from 100 g of the K salt of V with 2 N KMnO_4 (5 atoms O) at 100 prisms with 4 mols H_2O (0.5 mol is firmly held even at 135°), $[\alpha]_D^{20} -45.68^\circ$ (H_2O 1.345). IX, from 50 g VII allowed to stand 3 hrs at 37° in 250 cc. of 70% AcOH, : 120.5° , $[\alpha]_D^{20} -159.3^\circ$ (H_2O , c 1.639). Tri-K salt of II (164 g from 756 g of the K s of I heated 10 hrs at 100° with 1264 g KMnO_4 in 201 H_2O) $[\alpha]_D^{20} 20.35^\circ$ (H_2O , c 3.31.

C. A. R.
Lignin, humic acid and humin. W. FUCUS. *Z. anorg. Chem.* 44, 111-8 (1931).
cf. C. A. 25, 761.—Investigations by F. and co-workers are summarized. E. SCHÖTTE.
Lignin and related compounds. V. Action of halogens on lignin and woc.
HAROLD HIBBERT AND CHARLES A. SAKKEY. *Can. J. Research* 4, 110-8 (1931);
C. A. 24, 5151, 5153.—Spruce meal and a variety of lignin products were treated w
Br dissolved in a variety of solvents. The addn of Br was found to be a function
time, concn, solvent used and the acidity of the reacting medium, and quite unsat
factory for the detn of unsatn. VI. The mechanism of aqueous halogenation.
AUSTIN TAYLOR, O. MAASS AND HAROLD HIBBERT. *Ibid.* 119-33.—The rates of add
of HClO and HBrO to several unsatd compds, under the influence of various catalys
have been studied. The addn of HClO to allyl alc. and dipropenylglycol is catalyz

by both H^+ and Cl^- , the catalytic effects of these ions, mole for mole, being equal. T

effects of H^+ and Cl^- were similar, together w HCl w HBr the sum of their sep effects but
proportional to their product, indicating undissocd HCl as the real catalyst. Tl
catalytic effect of varying amts of added HCl is approx. proportional to the square
the amt. added, again indicating undissocd HCl as the real catalyst. The addn
 HBrO to allyl alc. is catalyzed by HCl and HBr to about the same degree. This cataly
of HBrO addn by HCl provides further support for the hypothesis of undissocd HCl
the real catalyst in the addn of HClO . This catalytic effect of undissocd halogen ac
on the addn. of hypohalous acids is given as an explanation of the greater rates of add
of HClO and HBrO from Cl and Br water as compared with solns of HClO and HBr .
The addn. of HBrO to fumaric acid is not catalyzed by HBr but, on the other hand,
considerably decreased. It is catalyzed, however, by HCl , although not nearly to tl
same degree as the addn. to allyl alc. This difference in behavior between fumaric ac
and allyl alc. shows the necessity for studying the effect of HCl and HBr on the rate
addn. of HBrO to a wide variety of unsatd compds, before any general theory of a
halogenation applicable to all types of unsatd compds can be put forward. Tl

possible influence of conjugation has, however, been pointed out. An attempt has been made to apply the theory of hypohalous acid addn. catalyzed by undissoc. halogen acid to the aq. halogenation of phenols. The reaction between HBrO and m -nitrophenol has been shown to be catalyzed by HCl although not to the same extent as by HBr . This indicates HBrO addn., catalyzed by undissoc. HBr , as one of the mechanisms involved in the aq. bromination of phenols. A similar mechanism cannot be the fastest one in the aq. chlorination of phenols, since it is in disagreement with kinetic data of Soper and Smith (*C. A.* 20, 2840), showing Cl to react faster than HClO with phenols. Resorcinol and phloroglucinol do not fall under the general classification proposed by Soper and Smith for phenols. Even here, however, the main reaction between Cl water and the phenol seems to involve Cl rather than HClO . Several parallel mechanisms may be involved in the aq. halogenation of phenols and the necessity for further work in this field has been emphasized.

J. W. SMITH

The enolization of ketones. V. GRIGNARD AND H. BLANCHON. *Bull. soc. chim.* (4), 49, 23-42 (1931). — The Grignard reagent has an enolizing reaction on ketones (*C. A.* 21, 731), and it is possible to esterify the mixt. of enol so produced and the accompanying tertiary alc. from the normal reaction by means of acid chlorides or anhydrides. After sepn. of the esters by distn., the enolic ester can be hydrolyzed with H_2O . With certain precautions the enol keto mixt. can be isolated by the decompn. of the Mg complex by NH_4Cl (*C. A.* 20, 751). A systematic study was undertaken to establish a relation between the constitution of the ketone, its aptitude for enolization and the stability of the enol, the present paper dealing with the enolization of ketones (*C. A.* 24, 1342), and the prepn. of the enols. To obtain the max. amount of enolate iso-PrMgBr was selected as being the most readily prepd. reagent. In general, the ketone was dissolved in an equal vol. of Et_2O and gradually introduced into the reagent. After refluxing for 30 min. and then cooling to -5° , the theoretical amt. of AcCl in 5 vols. of Et_2O was added dropwise with stirring and the resulting yellow mass was dropped into ice water contg. NaHCO_3 . The washed and dried Et_2O exts., freed from Et_2O , were distd. under reduced pressure, giving 3 fractions, the keto-enol mixt., the enol acetate and the acetate of the tertiary alc. Cyclohexanone-enol — 100 g. of cyclohexanone gave 5 g. of keto-enol mixt. contg. 30% enol, 25 g. of the enol ester and 60 g. of isopropylcyclohexanol, b_p 123°. The ester was hydrolyzed with 10% $(\text{CO}_2\text{H})_2$ yielding a 70% enol, n_D^{20} 1.4601. The tautomerization to cyclohexanone was complete in 10 hrs., the transformation being retarded by $(\text{CO}_2\text{H})_2$, but not by CaH_2O or $\alpha\text{-CaH}_2(\text{CO}_2)\text{O}$, and accelerated by mineral acids, alkalis and Br . Thuyone-enol — Thuyone formed mainly isopropylthujol acetate, b_p 125°, on treatment with iso-PrMgBr , and also some thuyone-enol mixt. contg. 89.5% enol, n_D^{20} 1.4540, which reverts to the ketone in 15 hrs. The ketonization is completed in 2 hrs. with NaHCO_3 and instantaneously with Br , demonstrating the difficulty of obtaining the pure enol and the inapplicability of the Kurt Meyer method. Carvone-enol — 200 g. of carvone, enolized by EtMgBr , gave ethylcarvicol, b_p 142°, and 15 g. of 98.5% carvone-enol, b_p 114-5°, d_4^{25} 0.957, n_D^{20} 1.5151, R_M 47.2, calcd. 46.3. After 23 hrs. the mixt. contains 1.5% enol, tautomerization being accelerated by light, acids and alkalis. With Br the ketonization is complete in 1 hr. at 15° and in 15 mins. at 40° . The Kurt Meyer method applied at 0° would give results 10% below those furnished by the detn. according to Job and Reich. Enol of menthyl oxide — Enolization of 100 g. of $\text{Me}_3\text{C-CH-COMe}$ with iso-PrMgBr yielded in addn. to 31 g. of 2,3,5-trimethyl-4-hexen-3-ol, b 172°, and 10 g. of enol acetate, 21 g. of a 70% enol b 134°, d_4^{25} 0.8810, n_D^{20} 1.4500, R_M 30.8, calcd., 30.18, which ketonized in 5 hrs. to 6.4% enol. Enols can be formed through the chlorohydrins. These are transformed into the acetic ester by AcOK and then treated as above. From pulegone was obtained 50% yield of a 98.2% enol, b_p 79-81°, n_D^{20} 1.48115, d_4^{25} 0.9130, R_M 47.39, calcd. 46.76. The 75% enol derived from MeBr through the chlorohydrin was a very unstable liquid, b 205°. Tautomerization was complete in 10 hrs. in the dark, 5 hrs. in light and was instantaneous in the presence of acids and alkalis. The ketonization was accelerated with Br . The enol did not give a color reaction with FeCl_3 nor did it form an ammoniacal enolate. C. R. A.

Bromomesitylene. LEE I. SMITH (*Org. Syntheses* II, 24-5 (1931)) — Details are given of the bromination of $\text{C}_6\text{H}_3\text{Me}_3$ in CCl_4 at $10-5^\circ$, the yield of $\text{C}_6\text{H}_2\text{Br}_2\text{Me}_3$ being 79-82%, traces of side-chain derivs. are removed by treatment with EtONa . The Br deriv. b_p 132°, b_p 139°, b_p 140°, b_p 157°.

C. J. WEST

Process for the preparation of mesitylene. V. IPAT'EV, B. DOLGOV AND I. VOLNOV. *Ber.* 63B, 3072-8 (1930). — Ipatev, Jr., in connection with his work on the displacement of metals of the 5th group from their chlorides (SbCl_3 , BiCl_3) by H in Me_2CO under pressure, noticed an extensive condensation of the solvent, irrespective of the salt used,

which occurred even when the solns were dild. with H_2O . This made it seem probable that the condensation to mesitylene (I) and higher condensation products depends not only on the dehydrating action of the HCl but also on a catalytic action exerted by it. The present authors were chiefly interested in the mechanism of the condensation to I, and since the use of H_2SO_4 under the same conditions resulted in a very deep-seated condensation with formation of S-contg products, they discarded this acid and used HCl. By choosing the mildest possible conditions and studying the effect of varying them they were able to det. the optimum conditions for the prepn. of I. Using the Ipat'ev high-pressure app., they found that the yield of I after 24 hrs at 175° (initial pressure 100 atm) rapidly increases with increasing concn. of HCl up to 5 vol. % of d. 1.19 acid and then slowly falls as the HCl concn. is increased to 50 vol. %, where higher, terpene-like condensation products appear. This explains the poor yields of I from Me_2CO satd with HCl. Keeping the concn. of HCl const. (5 vol. %) and varying the temp. gave a very rapid increase in yield (based on the Me_2CO not recovered) with increasing temp. up to 140° , the yield based on the total amt. of Me_2CO charged into the app. reached its max. at 170° ; from 175° to 190° the 2 curves coincided, pointing to a complete utilization of the Me_2CO charge. A very significant fact is the decrease in lower, O-contg. condensation products (mestyl oxide) with increasing temp. Under optimal conditions (5 vol. % and 145°), the yield of I no longer increases after 24 hrs., up to this point the yield increases very regularly. Under these conditions 3–4.5% mestyl oxide could be isolated as the primary condensation product. The absence of phorone in the condensation products indicates that either it is not formed at all under these conditions or quantitatively undergoes further condensation. Under the above optimal conditions, com. Me_2CO gives 43% I, based on the Me_2CO not recovered, up to 50% of the Me_2CO charged does not react but it can be put through the process again. At $175-80^\circ$ practically all the Me_2CO is used up, giving 36% I, as against only 17.5% previously obtained. C. A. R.

Isodurene (1,2,3,5-tetramethylbenzene). LEE I SMITH. *Org. Syntheses* XI, 66–9 (1931); cf. C. A. 24, 1850—2,4,6- $Me_3C_6H_2Br$ gives a Grignard reagent which yields 53–61% of 1,2,3,5 $C_6H_4Me_4$ with Me_2SO . C. J. WEST

The forced reaction between tetraphenylethylene and some organomagnesium derivatives. HENRY GILMAN and STANTON A. HARRIS. *Bull. soc. chim.* [4], 49, 10–12 (1931).—The formation of o-phenylbenzohydrylaniline by the action of $PhMgBr$ on $Ph_2C:NPh$ (C. A. 23, 3909), in which an addn. is made to a benzene ring in a compd. contg. several closely situated Ph groups, suggested the corresponding reaction with $(Ph_2C)_2(I)$. Fifteen g. of I (0.045 mol.) in 150 cc. of $PhMe$ was heated with 0.174 mol. of $PhMgBr$ at 60° for 20 hrs. and then at $90-100^\circ$ for 6 hrs., the mixt. being guarded from the atm. Eighty-six % of I was recovered and on addn. of CO_2 to the reaction mixt. before hydrolysis only $BzOH$ was formed. Similarly, after a forced reaction with $PhCH_2MgCl$, 85.5% of I was recovered and carbonation with CO_2 gave $PhCH_2CO_2H$ uniquely. Neither expt. showed any evidence of any reaction of the type expected in spite of the forced conditions and the highly phenylated nature of I. C. R. A.

2,4-Diaminotoluene. S. A. MAHOOD and P. V. L. SCHAFFNER. *Org. Syntheses* XI, 32–5 (1931).—Details are given of the reduction of 2,4-(O_2N) $_2C_6H_3Me$ with Fe and HCl, the di- NH_2 deriv. being obtained in 74% yield. C. J. WEST

p-Iodoaniline. R. Q. BREWSTER. *Org. Syntheses* XI, 62–4 (1931).— $PhNH_2$ in $NaHCO_3$ treated with 1 mol. powd. I_2 gives 75–84% of p- $IC_6H_4NH_2$, m. $62-3^\circ$. C. J. WEST

The action of ethyl nitrate on phenylhydrazine in the presence of sodium ethylate. EUGEN BAMBERGER and OTTO BILLETTER. *Helv. Chim. Acta* 14, 219–32 (1931).—The method used for the prepn. of $PhN:NOH$ from $PhNH_2$, $EtONO_2$ and $NaOEt$ (C. A. 15, 1516) was applied to $PhNHNH_2$ with the hope of obtaining $PhN(NO_2)NH$ but the attempt failed. $NaNO_2$, C_6H_5 , $PhNH_2$, $PhNHNHAc$, PhN_3 , $(PhN)_3$, $HOAc$, N_2 and $MeC(N:NPh).NNHPh$ (I) were isolated from the reaction mass. Mechanisms for the formation of these compds. are suggested and evidence in their support is offered. I was prepd. in 94% yield from $PhNHNHCHMe$ (II) and $PhN(:N)Cl$. In the same manner nitromethylformazyl, $MeC(NNHPh)N.NC_6H_4NO_2$, m. 154° (with 0.5 mol. of $EtOH$ of crystn.), was prepd. in quant. yield from II and p- $NO_2C_6H_4N(:N)Cl$. I decomps. in the presence of dil. and concd. HCl to yield $PhN(:N)Cl$ and $PhNHNH_2$, resp. I. M. LEVINE

Condensation of ethyl acetoacetate with aromatic amines. I. G. V. JADHAV. *J. Indian Chem. Soc.* 7, 669–76 (1930).— $AcCH_2CO_2Et$ (I) and aromatic amines give $MeC(NHR).CHCONHR$ and either $MeCOCH_2CONHR$ (II) or $MeC(NHR).CHCO_2Et$ (III), the effect of substituent groups in the aromatic nucleus was studied. Compds. of

type II were isolated in only a few cases and heating the reaction mixt was shown to convert them to compds of type III, except in the case of 1,1- β -*p*-nitroanilinoacetonate. The following compds were prepd by refluxing equal wts of I and the proper amine unless otherwise indicated: *acetoacet m*-nitroanilide (IV), pale yellow, m 120-1°, β -*m*-nitroanilinoaceton m-nitroanilide (V), yellow, m 153-3°, 1,1- β -*p*-nitroanilinoacetonate, yellow, m 122-3°, β -*p*-nitroanilinoaceton p-nitroanilide (VI), orange, m above 250°, *acetoacet p*-anisilide (VII), m 117-8°, β -*p*-anisidinocroton *p*-anisilide, m 235-6°; Et *p*-phenetidinocrotonate m 60-1° (from a cold mixt on standing 48 hrs), β -*p*-phenetidinocroton *p*-phenetilide m 230-1°, β -*m* (1,3,5) xylidinocroton m (1,3,5) xylidide, m above 275°, β -*m* (1,3,4) xylidinocroton m (1,3,4) xylidide m above 275°, *acetoacet-m*-chloroanilide, m 105-6°, β -*m*-chloroanilinoaceton *m*-chloroanilide m 210 1°, *acetoacet p*-chloroanilide (VIII), m 132-3°, Et β -*p*-chloroanilinoacetonate, m 55° (from a cold mixt. on standing 72 hrs) β -*p*-chloroanilinoaceton *p*-chloroanilide, m above 275°, *acetoacet-o*-chloroanilide (IX), m 107-8°, and β -*o*-chloroanilinoaceton *o*-chloroanilide, m 236°. V and VI with HCl give IV and *acetoacet p*-nitroanilide (X), m 123-3°, resp. IV, X, VIII, IX and VII with NH_4OH give the following anilides of β -aminoacetic acid *m*-NO₂, orange red, m 129-30°, *p*-NO₂, yellow m 189-90°, *p*-MeO, m 109-10°; *p*-Cl, m 110°, and *o*-Cl, m 96-7°.

A. F. SHEPARD

Azobenzene. H. L. BIGELOW AND ALBERT PALMER. *Org Syntheses* XI, 16-8 (1931)—Details are given of the action of As_2O_3 and NaOH upon PhNO , $\text{PhN}(\text{O})\text{Ph}$ results in 85% yields, the compd is easily volatile with steam at 140-50°.

C. J. WEST

Remarks concerning the most recent work of A. Angeli. EUGEN BAMBERGER. *Helv Chim Acta* 14, 242-9(1931)—A mechanism offered by B for the reaction by which azohydroxyamides $\text{RN}(\text{OH})\text{N}(\text{NR})_2$ are obtained from the interaction between PhNO (I) and PhNHNH_2 (II) or their derivs involves an aldol condensation of the reactants followed by a loss of H_2 . This mechanism is in contrast to that given by Angeli who supposed that II is first reduced to $\text{PhN}(\text{H})\text{NH}_2$ which then adds L. Several instances of reactions easily explained by means of the aldol condensation are given. Besides the simple compds, β - $\text{BrC}_6\text{H}_4\text{NO}$ (III), PhMeNNH_2 (IV) and Ph_2NNH_2 (V) undergo the same type of reaction. From IV and I, phenylazomethylphenylidoxime, $\text{PhMeNN}(\text{O})\text{Ph}$ (VI), m 72°, $\text{PhMeNN}(\text{NPh})$ (VII) and PhNHMe (VIII) were obtained. By reduction of VI with Zn, VII, VIII, PhNH_2 , and NH_3 were produced. Condensation of V and I yields phenylazodiphenylamidoxime, $\text{Ph}_2\text{NN}(\text{O})\text{Ph}$, m 128.5-9°, Ph_2NH and PhNHOH . From III and IV β -bromophenylazomethylphenylidoxime, $\text{BrC}_6\text{H}_4\text{N}(\text{O})\text{N}(\text{NMePh})$ (IX), m 77-8°, $\text{BrC}_6\text{H}_4\text{N}(\text{O})\text{C}_6\text{H}_4\text{Br}$ (X), and possibly $\text{BrC}_6\text{H}_4\text{N}(\text{NMePh})$ were isolated. Reduction of IX yielded $\text{BrC}_6\text{H}_4\text{NH}_2$ (XI) and VIII. III and V interact to give β -bromophenylazodiphenylamidoxime, $\text{BrC}_6\text{H}_4\text{N}(\text{O})\text{NNPh}_2$ (XII), m 119-20°, X and $\text{BrC}_6\text{H}_4\text{NH}_2$. The reduction products of XII consisted of X, $\text{BrC}_6\text{H}_4\text{NH}_2$, XI, Ph_2NH , $\text{BrC}_6\text{H}_4\text{NH}_2$, and Ph_2NHAc .

I. M. LEVINE

***N*-Diethylamino-*N*-phenylurea.** R. STOLLÉ AND W. BRANDT. *J prakt Chem* 129, 206(1931)—Hurd and Spence (C. A. 21, 570) state that Et_2NNH_2 and PhNCO give the compd $\text{PhN}(\text{C}(\text{NEt}_2)_2)_2$, S and B state the compd analyzes for $\text{C}_{11}\text{H}_{15}\text{ON}_2$ and is thus a semicarbazide deriv.

C. J. WEST

***p*-Methoxybenzylmagnesium bromide.** HENRY GILMAN AND E. A. ZOELLNER. *Bull soc chim* [4], 49, 7-9(1931)—The extraordinary activity of the halogen in the halides of the type $\text{p-MeOC}_6\text{H}_4\text{CH}_2\text{X}$ renders these compds so unstable that various workers have reported that $\text{p-MeOC}_6\text{H}_4\text{CH}_2\text{Cl}$ does not form a Grignard reagent (C. A. 18, 237, 23, 823). By a modification of the usual procedure (C. A. 23, 2934) $\text{p-MeOC}_6\text{H}_4\text{CH}_2\text{MgBr}$ has been prepd in satisfactory yields. To 1.82 g (0.075 atom) of Mg, 30-80 mesh, covered with 11.5 cc Et_2O , were added 10 drops of freshly distd $\text{p-MeOC}_6\text{H}_4\text{CH}_2\text{Br}$ (b_p 129°, d_4 1.41), and then 0.05 g of I. After boiling for 10 mins without stirring, 4.35 g (0.0216 mol) of the bromide mixed with 23.5 cc. of Et_2O , was added uniformly over a period of 30 mins with stirring and refluxing. After the addn. this treatment was continued for 10 mins. The resulting soln. gave a positive reaction for organometallic reagent (C. A. 19, 2473) and titrated for 54% $\text{p-MeOC}_6\text{H}_4\text{CH}_2\text{MgBr}$ (C. A. 17, 530). Treatment with CO_2 gave $\text{p-MeOC}_6\text{H}_4\text{CH}_2\text{CO}_2\text{H}$, m 84-5°, and some $(\text{p-MeOC}_6\text{H}_4\text{CH}_2)_2$. The yield could be raised to 75% by using 10-20 atoms of Mg and about 30 mols of Et_2O per mol of bromide, an even better yield should be obtained of the corresponding organomagnesium chloride.

C. R. ADDINALL

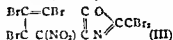
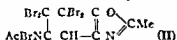
Sulfur derivatives of 2-methoxytoluene. G. B. KOLBATEKAR AND K. V. BOKIL. *J Indian Chem Soc* 7, 843-50(1930)—3,4-Me(MeO) $\text{C}_6\text{H}_3\text{SO}_2\text{Cl}$ (I) and 2-MeOC $\text{C}_6\text{H}_4\text{Me}$ in CS_2 with AlCl_3 give 3,3'-dimethyl-4,4'-dimethoxydiphenyl sulfone (II), m. 138°. Alk. KMnO_4 oxidizes II to 3,3'-dicarboxy-4,4'-dimethoxydiphenyl sulfone (III), sinters 245°.

dark red color. 4,4'-(1 Nitrobenzoyl)diphenyl ether, m 175°, sol in concd. H_2SO_4 with a yellow color, ds 3 NH_3 deriv., greenish yellow, m 150-1°; H_2SO_4 soln., red-orange; azo dye from β -C₆H₄OH, orange, sol in H_2SO_4 with a red color. 3 (3 Nitrobenzoyl)diphenyl sulfide m 128-0°, sol in concd H_2SO_4 with an orange red color; the di-4,4'-(3 nitrobenzoyl) deriv. m 221-30° and gives an orange red color in H_2SO_4 . Di-4,4'-(4-nitrobenzoyl) deriv., m 278°. H_2SO_4 gives an orange red color. Di-4,4'-(3 nitrobenzoyl)diphenyl selenide, greenish yellow m 221-2°, sol in H_2SO_4 with an orange red color; the di-4,4'-(4-nitrobenzoyl) deriv. yellow m 267-8° and gives a deep red soln. in H_2SO_4 . 4-Benzoyl-4'-nitrodiphenyl sulfide, m 145°, concd H_2SO_4 soln., orange. C. J. WEST

Aromatic disulfides and Suggen's parachors. S. S. BHATTAGAR AND BALWANT SINGH. *J. Indian Chem. Soc.* 7, 663-6 (1930).—The parachor values for several aromatic disulfides indicate that the conventional structure with 2 bivalent S atoms is correct. The following disulfides were studied and data for each were detd at 3 or 4 different temps: diphenyl, m 66.5°, $d_4^{18} = 1.110$, γ 35.08, parachor (mean) 477.9; 4,4'-dichlorodiphenyl, m 71.5°, $d_4^{12.5} = 1.204$, γ 39.28, parachor (mean) 550.6; 4,4'-dibromodiphenyl, m 93.8°, $d_4^{11} = 1.617$, γ 40.23, parachor (mean) 575.9; 4,4'-dimethyldiphenyl (I), m 46.0°, $d_4^{11.5} = 1.068$, γ 33.13, parachor (mean) 552.0; 4,4'-dimethoxydiphenyl (II), m 73.5°, $d_4^{10.5} = 1.139$, γ 37.84, parachor (mean) 605.1; β , β '-dinaphthyl, m 139.0°, $d_4^{11} = 1.144$, γ 37.20, parachor (mean) 659.6; dibenzyl, m 71.0°, $d_4^{10.5} = 1.085$, γ 36.42, parachor (mean) 557.1; 2,2'-dinitrodiphenyl, m 112.0°, $d_4^{11} = 1.308$, γ 44.81, parachor (mean) 655.7. The parachor value of I is 6.3 units lower than that calcd. for bivalent S, presumably this is due to assocn. A. T. SHEPARD

Action of bromine on phenols. GUSTAV JELLER, WERNER DIETRICH, THEODOR HÄRMER, HORST KÄTZEL, LUDWIG ROTTSAHL AND P. G. ZAMBALOS. *J. prakt. Chem.* 129, 211-56 (1931).—The action of Br upon phenols may be divided into several categories. (1) Br does not react.—This is the case where the phenol group is protected as in 4,6,3-Br₃(AcNH)C₆H₃OH and 2,4,6,3,5-Br₅(AcNH)(O₂N)₂C₆H₂OH is also unchanged on heating with Br-AcOH. (2) Br enters only in the nucleus.—4,6,3-(O₂N)₃(AcNH)C₆H₃OH and Br in AcOH, heated 8 hrs. on the H_2O bath, give 4,6-dinitro-3-amino-2-bromophenol, yellow, m. 207°; di-Ac deriv., m. 222°; partial hydrolysis with dil. NaOH gives the 3-AcNH deriv., m. 223.5°. Warming on the H_2O bath with concd. HNO_3 gives 4,6-dinitro-2-bromo-1,3-dihydroxybenzene, golden yellow, m. 102.5-3°, 8% NaOH gives a carmine-red salt; this also results by diazotizing in concd. H_2SO_4 . 2-Acetamido-4-nitroresorcinol and Br AcOH, shaken until soln. results and allowed to stand 24 hrs., give the 6-Br deriv., pale yellow, m. 173-4°. 6,3-O₂N(AcNH)C₆H₃OH and Br-AcOH, warmed on the H_2O bath, give 2,4-dibromo-6-nitro-3-aminophenol, yellow, m. 199-200°; the alkali salt is orange-yellow, and 2-bromo-6-nitro-3-aminophenol, yellow-brown, m. 230°; the structure of these compds was established by removal of the NH₂ group. 3,5,4-(O₂N)₃(AcNH)C₆H₃OH and Br-AcOH give the 2,6-di Br deriv., decomps 274.5°; the free NH₂ compd., brown, m. 138°. (3) In addn. to substitution some of the groups are replaced by Br.—4,3,5-H₂N(O₂N)₃C₆H₃OH and Br AcOH, warmed 2 hrs. on the H_2O bath, give bromoanil. Energetic bromination of 4,3-O₂N(AcNH)C₆H₃OH gives 2,3,4,6-Br₄C₆H₂OH, with 1 mol. Br in AcOH there results the 6-Br deriv., yellow, m. 205° (decomps). EtOH-KOH gives 6-bromo-4-nitro-3-aminophenol, red brown, m. 244° (decomps), whose structure was established by converting it into 6,4-Br(O₂N)C₆H₃OH, m. 113°. With more than 1 mol. Br, there also results the 2,6-Br₂ deriv., yellow-orange, m. 205°; 2,6-dibromo-4-nitro-3-aminophenol, egg-yellow, m. 162.5°. 2,4-Dinitroresorcinol and a slight excess of Br in AcOH give the 6-Br deriv., m. 69°, a small quantity of the 5-Br deriv. and tetra-Br deriv. 4-Acetamido-6-nitroresorcinol and Br give the 2-Br deriv., m. 226° (decomps). 3- and 4-Nitropyrocatechols, on energetic bromination, give the tetra Br deriv. and tetrabromo- α -quinone. (4) The bromination is complete and leads to Br₄C₆H₂OH.—This is true of 3,4-O₂N(AcNH)C₆H₃OH; 3,4-O₂N(H₂N)C₆H₃OH, on the other hand, gives bromoanil. (5) Besides complete bromination there follows addn. of Br and entrance of a halogen atom in the AcNH group.—2,4-(AcNH)₂C₆H₃OH and 1 mol. Br in AcOH give the 6-Br deriv., m. 215°, with HNO_3 (d. 1.4) and H_2O (1.4) on the water bath, there results 2-acetamido-6-bromoguanone, orange, m. 183°; SO_2 in EtOH gives the hydroquinone, m. 154°, excess Br gives tetrabromodiketone-R-pentene (I), m. 143-4°. Two mols. Br gives 2,4-diacetamido-5,6-dibromophenol, cryst. with 2 mols. H_2O , m. 188.5°, anhyd., it m. 208°. HNO_3 gives 2-acetamido-5,6-dibromoguanone, orange, m. 213°; the hydroquinone m. 188°. Four mols. Br gives 2-acetamido-3,5,6-tribromoguanone, golden yellow, m. 128°. Excess Br gives 4-(N-bromoacetamido)-5,6-tetrabromo-2,1-ethenylaminophenol (II), light orange, m. gradually above 210°, and I;

with 2 *N* NaOH there results 4-acetamido-5,6-dibromo-2,1-ethenylaminophenol, m. 245°; fuming HCl gives 5,6,2,4-Br₂(AcNH)₂C₆H₂OH, m 183.5° 2,6-(AcNH)₂C₆H₂OH and Br in AcOH give the 1,5(?)—Br₂ deriv, m 209°, a mono-Br deriv, m 215° (decompn). Excess Br gives an orange-yellow addn product, amorphous, m 181-3° (decompn), which gives with AcONa 6-acetyl-amino 1,4,5-tribromo-2,1-ethenylaminophenol, m. 226°.



Oxidation of 2,6-(AcNH)₂C₆H₂OH with HNO₂ gives 2,6-diacetamidohydroquinone, orange, m 270° (decompn). PhNH₂ gives an anilide, blue black, m 202°, the same quinone results from 2,4,6-(AcNH)₂C₆H₂OH. (6) A quinone is formed with the retention of the AcNH group—2,4-(AcNH)₂C₆H₂OH, NaNO₂ and dil H₂SO₄ give 2,4-diacetamido-6-nitrophenol, orange, m 215°, Br in AcOH gives 2-acetamido-3,5,6-tribromohydroquinone. (7) In the case of quinone formation there generally follows complete bromination with the removal of substituents and formation of bromoanil—3,2,4-H₂N(O₂N)₂C₆H₂OH and Br in AcOH, warmed 4 hrs on the water bath, give the 6-Br deriv yellow m 148° (di Ac deriv, m 141°, mono Ac deriv pale yellow, decompn 221°) excess Br gives bromoanil. Through the diazo reaction in concd H₂SO₄ there results 2,4-dinitro-6-bromoresorcinol, m 89° HNO₂ (d 1.42) gives typhonic acid 4,2,6-AcNH(O₂N)₂C₆H₂OH and excess of Br give bromoanil 2,5-Dinitrohydroquinone gives the same product. (8) Special behavior of 4,2,1,3-H₂N(O₂N)₂C₆H₂OH—Br in AcOH gives 5-bromo-2-nitro-1,4-quinone azide, yellow-brown, m 186° (decompn), the Ac deriv gives bromoanil 2-Acetamido-3,5,6-tribromophenol, light yellow m 151°, Br gives a compd, m 218° (decompn), the free NH₂ compd with Br gives yellow needles, exploding at 180°. (9) In several cases the Br opens the C₆H₄ ring with the formation of tetra-romodiketo *R* pentenes—4-Nitro-2-acetamidophenol, decompn 267°, and Br in AcOH give the 6-Br deriv, m 206°, this also results by brominating 2,4-(O₂N)₂C₆H₂OH, reducing and acetylating. Excess of Br in AcOH gives 1 and 3,5,6-tribromo-2-amino-1-quinone 4-dibromide(?), m 216° (Ac deriv, m 203°). Heating 2,4,6-AcNH(O₂N)₂C₆H₂OH with excess Br several hrs also gives 1. 2-Acetamido-5-nitrophenol, m 234-9° 2,3,5,6-H₂N(O₂N)₂C₆H₂OH and HNO₂ give 3,5,6-trinitro-1,2-quinone azide, yellow, sepg with C₆H₅ of crystn and is very explosive, the alk soln gives a deep violet red color with β C₆H₅OH, cooling with 4-8 parts LiOH gives 2,3,5-(O₂N)₃C₆H₂OH; a shorter period of heating gives 3(?)ethoxy-5,6-dinitro-1,2-quinone azide, yellow, m 166° (decompn), alk β C₆H₅OH gives a deep brick red dye. 2,5-Acetamidonitrohydroquinone and Br-AcOH give 2-acetamido-5,6-dibromohydroquinone, orange yellow, m 213°, identical with that obtained on oxidation of 5,6,2,4-Br₂(AcNH)₂C₆H₂OH; an intermediate product is 2-acetamidotribromohydroquinone, yellow, m 183-90° (decompn). Hexaacetyl-2,5-diaminophenylol, (AcNH)₂C₆(OAc)₄, m 216° (decompn), 5 mols Br give 2,5-diacetamidohydroxyquinone, red-brown, does not m 240°. 3,4-(AcNH)₂C₆H₂OH and Br-AcOH give the 6-Br deriv, decompn 238° and the 2,6-Br₂ deriv, m 223°, excess Br gives 2-methyl-5,7-dibromo-6-hydroxybenzimidazole, greenish flakes which do not crystallize. 2,6-Diacetamidohydroquinone, m 210° (decompn); Br gives 3-bromo-2,6-diacetamidohydroquinone, red brown, m 225°, SO₂ in AcOH at 40-50° gives 3-bromo-2,6-diacetamidohydroquinone, m 199° (decompn). With 6 mols Br there results 3,5-dibromo-2,6-diacetamidohydroquinone, red-brown, m 201°, the mother liquor contains a small quantity of (CBr₃)₂CO, I could not be detected. (10) Formation of hexabromoacetone: Excess of Br-AcOH with 2,6-diacetamidohydroquinone gives (CBr₃)₂CO, this also results from nitroaminophenylol, in addn to Br₂C₆(OH)₄, a compd C₆H₂O₂NBr₂ was obtained in some expts. Pentaacetyl-nitroaminophenylol, m 191°, is formed with Ac₂O and AcONa, in the presence of H₂O there results the di Ac deriv, light yellow, explodes at 214°. 2,3-O₂N(H₂N)₂C₆H₂OH and 1 mol Br in AcOH give the 6-Br deriv, m 233°; excess Br gives the 4,6-di Br deriv, yellow, m 182°, warming with 6 mols Br for 16 hrs. gives a mixt. of 2-dibromomethyl-4-nitro-5,6,7-tribromobenzoazole, m 233°, and the 2-tribromomethyl deriv. (III), m 172°. Tetra-bromodiketopentene and PhNH₂ in EtOH give 4,5-dibromo-2-dianilino-3-anil-1-ketopentamethylene, m 261°.

C J WEST

Quinone formation from nitroacetamidohydroquinone. GUSTAV HELLER AND THEODOR HELLER. *J. prakt. Chem.* 129, 207-10(1931)—Dinitrohydroquinone is reduced by SnCl₂ and HCl in AcOH to 2,5-nitroacetamidohydroquinone (I), brick-red, decompn 226°; the tri-Ac deriv m. 183-4°. 2,5-Nitroaminohydroquinone, red, carbonizes 151°; heating with dil alkali gives a color change from blue through green to yellow; no characteristic substance was obtained through diazotizing. I and fuming HNO₂ in AcOH give 2-nitro-5-acetamido-1,6-dihydroxyquinone, orange-yellow, m 164°.

(decompn) crystals with 0.5 mol H_2O , with $PhNH_2$, there results 3,6-dianilino-2-acetamido-5-nitroquinone (II), steel blue needles, gradually decompn at 200° . Careful oxidation of I with CrO_3 in $AcOH$ gives apparently a monohydroxynitroacetamidoquinone, gradually darkens above 140° , $PhNH_2$ gives II. C. J. West

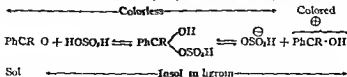
Halogenation of optically active phenylmethylcarbinol in the presence and in the absence of pyridine by thionyl chloride and the chlorides and oxychloride of phosphorus. JOSEPH KRYVON, HENRY PHILLIPS AND FRANK M. H. TAYLOR. *J. Chem. Soc.* 1931, 382-9.—The halogenation of 1- $PhMeCHOH$ by $SOCl_2$, PCl_3 , PCl_5 and $POCl_3$ in the presence and absence of C_6H_5N is studied. It is found that the presence of either pyridine or quinoline causes a reversal in the sign of rotation of the $PhMeCHCl$ when $SOCl_2$ is used. When PCl_3 , PCl_5 or $POCl_3$ are used in the presence of C_6H_5N , the signs of rotation of the $PhMeCHCl$ obtained are the same as in its absence but the magnitudes of the rotations are greater. The presence of K_2CO_3 is without influence on the magnitude of the rotatory power of the $PhMeCHCl$ obtained by the interaction of the 1 alc with PCl_3 . It is probable that the HCl produced in the initial stage of this interaction plays no further part in the reaction and that the $PhMeCHCl$ produced arises through the decompn of the intermediate compd, $PhMeC(ORCl)H$. In the absence of C_6H_5N , which causes the inversion reaction to predominate, the decompn of the intermediate compd appears to follow an indefinite course in which all the methods of decompn may occur, since a $PhMeCHCl$ with $[\alpha]_{D20} 10.1^\circ$ is obtained in place of the probably optically pure chloride with $[\alpha]_{D20} 64.1^\circ$ which is obtained when the decompn occurs in the presence of C_6H_5N . It would appear that, in the absence of C_6H_5N , the decompn of the intermediate compd $PhMeC(ORCl)H$ is initiated by the sepn of a phenylmethylcarbonium cation without the simultaneous production of a Cl anion. The results of the various expts. may be tabulated as follows:

Halogenating agent		Tert. base	$[\alpha]_{D20}$ of 1 alc	$[\alpha]_{D20}$ of $PhMeCHCl$
$SOCl_2$	2.5 mols	None	-51.7°	-63.5°
$SOCl_2$	2.0 mols.	C_6H_5N 1 mol	49.8	17.3
$SOCl_2$	2.0 mols.	C_6H_5N 2 mols	49.8	30.3
$SOCl_2$	2.0 mols.	Quinoline 2 mols	43.7	16.1
PCl_3	3.5 mols	None	51.7	10.1
PCl_3	2.0 mols	K_2CO_3	51.7	10.0
PCl_3	2.0 mols.	C_6H_5N 1 mol	49.8	64.1
$POCl_3$	2.5 mols	None	28.0	0.5
$POCl_3$	2.0 mols	C_6H_5N 1 mol	52.1	12.6
PCl_5	2.0 mols	None	43.7	4.5
PCl_5	2.0 mols	C_6H_5N 1 mol	49.8	10.4

The b , p , and n are given for many of the products.

C. J. West

Salt-forming characteristics of doubly- and singly-linked elements of the oxygen groups. I. The carbonyl group in benzaldehyde and acetophenone. JOHN W. BAKER. *J. Chem. Soc.* 1931, 307-14.—The investigation of the salt-forming characteristics of the carbonyl O in derivs of the type $ArCOR$ has been carried out along 2 lines. The first involves investigations of the condition of such groups in H_2SO_4 soln by the purely phys methods of partition, colorimetry, absorption spectra, etc. Secondly, the conversion of a pseudo-base system, directly attached to a C_6H_5 nucleus, into the "onium" salt will affect the directive action of the group causing enhanced tendency toward m -substitution. Pure BzH was distributed, under comparable conditions, between H_2SO_4 (of various concns) and ligroin, both with and without the addn of $(NH_4)_2SO_4$ to the acid layer. The following results were obtained $(100[H_2SO_4]/([H_2SO_4] + [H_2O]))$ and proportion of BzH in ligroin layer (%) with and without $(NH_4)_2SO_4$ are given: 7.3% free SO_3 , 0.5, 1.4, 10.9, 0.81, 1.03, 9.0, 0.83, 8.7, 8.0, 1.77, 30.8. The equil conditions in H_2SO_4 may be represented by the scheme:



In the absence of $(NH_4)_2SO_4$, these equil are displaced largely toward the right, the BzH being retained almost completely in the acid layer, even in 80% H_2SO_4 . Addn

oil whose *HCl* salt is bright red. PhMeC NNHPh and $p\text{-ONC}_6\text{H}_4\text{N}_2\text{Cl}$ in EtOH give acetophenone *p*-nitrophenylhydrazones brown *m* 194°. $p\text{-BrC}_6\text{H}_4\text{NNHNCPhMe}$ and PhN_2Cl give the same product. II and $p\text{-ONC}_6\text{H}_4\text{N}_2\text{Cl}$ give acetophenone 2 *p*-nitrobenzoyl-2,4-nitrophenylhydrazones. Cultrone, *m* 300°. III and $p\text{-ONC}_6\text{H}_4\text{N}_2\text{Cl}$ give the *p*-nitrophenylphenyl, benzylhydrazones $\text{PhMeC NNH(CH}_2\text{Ph)(CH}_2\text{NHCN}_2\text{ONC}_6\text{H}_4\text{NO}_2\text{)}$ *p* brown red *m* 130°. PhMeC NNHPh and $m\text{-(O-N)}_2\text{C}_6\text{H}_3\text{N}_2\text{Cl}$ in AcOH give acetophenone *m*-disitrophenylhydrazones, dark brown, *m* 234°. $\text{Ph}_2\text{C NNHPh}$ and PhN_2Cl do not react. $p\text{-O}_2\text{NC}_6\text{H}_4\text{N}_2\text{Cl}$ gives 30% of the *p*-nitrobenzoylphenylhydrazones, $\text{Ph}_2\text{C N-NHCH}_2\text{N}_2\text{N}_2\text{C}_6\text{H}_4\text{NO}_2$ *p* bluish red with violet metallic luster, *m* 194-6°. $\text{Ph(PhCH}_2\text{)C NNHPh}$ and PhN_2Cl give the compd. $\text{Ph(PhCH}_2\text{)C NNHCH}_2\text{N}_2\text{N}_2\text{Ph}$, golden yellow, *m* 135° (poor yield). *desoxybenzoin p*-nitrophenylhydrazones, red with bluish violet surface luster *m* 160°, results from $\text{Ph(PhCH}_2\text{)C NNHPh}$ and $p\text{-O}_2\text{NC}_6\text{H}_4\text{N}_2\text{Cl}$ or from $\text{Ph(PhCH}_2\text{)CO}$ and $p\text{-ONC}_6\text{H}_4\text{NNHPh}$. C. J. WEST

The condensation of phenolic aldehydes and their ethers with methyl propyl or butyl ketones. KUNJI IWAMOTO AND TARO KATO. *Science Rep's Tokyo Imp. Univ.* 1st Ser., 19, 684-93(1930) cf. C. A. 22, 575—Aromatic aldehydes with an OH or OMe group in the *p*-position condense with MeCOPr or MeCOBu to give substituted styryl Me ketones while *m* substituted aldehydes yield substituted distyryl ketones. *p*- $\text{HOC}_6\text{H}_4\text{CHO}$ (2.4 g), MeCOPr (1.9 g) and dry HCl give 2.9 g of α -ethyl-*p*-hydroxystyryl Me ketone (I), faint yellow, *m* 120-1°. The following styryl Me ketones were prep'd similarly: α -propyl *p*-hydroxy, pale yellow, *m* 100° (oxime, *m* 119.5-21.5°), α -ethyl *p*-methoxy (II), faint yellow, *b*₁ 171-2° (oxime, *m* 93-4.5°), α -propyl *p*-methoxy (III), pale yellow, *b*₁ 181-4° (oxime *m* 191-2°). I, II and III give the CHI_3 reaction. The distyryl ketones which follow were prep'd in the same way except that conc'd HCl was used in forming IV and V: α -ethyl *m*-*m'*-dihydroxy (IV), brick red, α -propyl *m*-*m'*-dihydroxy (V) brick red, α -ethyl *m*,*m'*-dimethoxy, dark yellow, and α -propyl *m*,*m'*-dimethoxy, dark yellow. A. F. SHEPARD

Behavior of stereoisomeric oximes of α,β -unsaturated ketones on hydrogenation. K. W. MEERZ. *Ber.* 61B, 2931-3(1930)—1-*o*-vanillacetone and $\text{NH}_2\text{OH HCl}$ in a little MeOH give 75% (as the *HCl* salt, needles, *m* 174° (decomp.)) of the *syn*-oxime (1 β' -hydroxy-4' methoxyphenyl 1-buten-3-one oxime), *m* 168-9°, which, heated 1 hr in 12 parts AcOH on the H_2O bath, changes into the *anti* form, *m* 142-3°. Both forms give by the Schotten-Baumann method the same *di* *Bz* deriv., *m* 149-9°, hydrolyzed by boiling alc. NaOH to the α,β -oxime. The latter is not hydrogenated in alc. contg. a little AcOH by H and Pd -charcoal, but the *syn*-oxime (4.5 g) yields a mixt. of 1.9 g $\text{HO(MeO)C}_6\text{H}_3\text{CH}_2\text{CH}_2\text{COMe}$ 0.9 g of its oxime and 1.6 g of $\text{HO(MeO)C}_6\text{H}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{N(H)Me}$. The same difference in the behavior of the *syn*- and *anti* forms on catalytic hydrogenation was noted with PhCH CHC(ONH)Ph and dypnone oxime, only the *syn* forms being hydrogenated and yielding the corresponding α,β -diamine. It is thus shown in 3 cases that it is possible by hydrogenation to det. to which series the oximes of α,β -unsat'd ketones belong. C. A. R.

Some reactions of phenyl propenyl ketone with semicarbazides and thiosemicarbazides. ALFRED Y. LIVINGSTONE AND FORSYTH J. WILSON. *J. Chem. Soc.* 1931, 335-7— PhCOCH CHMe yields a β -phenylsemicarbazone, *m* 212°, even with excess of reagent. The β -phenylthiosemicarbazone *m* 140°, becomes yellow in light without change in *m* $p\text{-NH}_2\text{NHCN}_2\text{NH}_2$ and the ketone in $\text{H}_2\text{O-EtOH}$ give thiosemicarbazidopropyl *Ph* ketone, *m* 140°. excess of reagent gives the same product. NH_2OH gives the oxime, $\text{C}_9\text{H}_9\text{ONS}$ *m* 165°. $p\text{-MeOC}_6\text{H}_4\text{COCH CHMe}$ gives a β -phenylsemicarbazone, *m* 249°. C. J. WEST

Some synthetic bases similar in ephedrine. E. CHERBULEZ, FR. NEUMEIER AND H. LOZERON. *Helv. Chim. Acta* 14, 196-200(1931)—Hydroxy and alkoxy derivs. of ephedrine (I) were prep'd to test their vasoconstrictor action in comparison with that of I. The method of prep'n. consisted of (a) the treatment of $\text{PhCH CHCH}_2\text{OH}$ (II) and its *Ac* (III) and *Bz* (IV) derivs. and alkyl ethers, with I_2 and HgO , in Et_2O and (b) subjection of the resulting iodohydrin (without sepn.) to the action of MeNH_2 under pressure, at 100°. The first step is favored by light. II, III and IV (the last 2 prep'd in 70 and 60% yields resp. by simple acylation of II) gave *di* α -hydroxy-ephedrine, $\text{PhCH(OH)CH(NHMe)CH}_2\text{OH}$ (V), *m* 110° (*HCl* salt, *m* 159.5°). Thus step (b) involved a hydrolysis. In the prep'n. of the alkoxy derivs. it was important to sep. the cinnamyl ethers from unchanged II, this was accomplished by benzoylation of the crude product, followed by fractionation. The intermediary compds., cinnamyl *Pr* ether, *b*₁ 131.2°, cinnamyl *Bz* ether, *b*₁ 135°, and cinnamyl allyl ether, *b*₁ 135°, were prep'd by the action of the alkyl iodide on the $\Delta\alpha$ deriv. of II (from II and NaNH_2 in C_6H_6) in 35-49%, 30-35% and 50-60% yields, resp., the *Me* and *Et* derivs. (*Et*, *b*₁

125–6°) were best obtained by means of the corresponding alkyl sulfates. The alkoxy derivs of I obtained were *methoxyephedrine* (VI) $\text{PhCH(OH)CH(NHMe)CH}_2\text{OMe}$, b_p 148° (yield 35%) (HCl , m 151–3°), *ethoxyephedrine*, b_p 148°, *ephedryl Pr ether*, b_p 151–3°, *ephedryl Bu ether*, b_p 155–7°, and *ephedryl allyl ether*, b_p 155–7°, all probably mixts of racemates. When the reaction was conducted in MeOH instead of Et₂O, II yielded 31% of 3,3-phenylmethoxy-2-methylaminopropyl alc., $\text{PhCH(OMe)CH(NHMe)CH}_2\text{OH}$ (VII), b_p 143°. VI was transformed into $\text{PhCH(OMe)CH(NHMe)CH}_2\text{OMe}$ by interaction with MeI. Of the derivs VI was the most active physiologically but it was much less active than I; in concns of 30 mg or more, per kg (for a cat), its action was reversed—it lowered the blood pressure to an extent increasing with increase in concn. V had a very light constrictor action while VII produced a reversal.

I. M. LEVINE

β -Phenylethylamines. I. Mescaline and mescaline-like substances. K. II. SLOTTA and H. HELLER. *Ber* 63B, 3029–44 (1930).—In addn to the sympathomimetic action which is a general property of β -aryl-ethylamines and -ethanolamines, mescaline, 3,4,5-(MeO)₃C₆H₂CH₂CH₂NH₂ (I), possesses a peculiar narcotic action. For a pharmacol and clinical investigation of I and a study of its relationships to other sympathomimetic drugs it was necessary to find a way by which it would be possible to prep. several hundred g of I. The only method by which it had hitherto been synthesized ($\text{RCHO} + \text{MeNO}_2 \rightarrow \text{RCH}_2\text{CHNO}_2 (+ \text{H}_2) \rightarrow \text{RCH}_2\text{CH}_2\text{NOH} (+ \text{H}_2) \rightarrow \text{RCH}_2\text{CH}_2\text{NH}_2$) serves to establish its structure rather than to prep the compd. The new method should also make it possible to prep. more easily the isomers of I, as well as mono- and dimethoxyphenylethylamines, in order to det. the influence of the position and no. of the MeO groups on the physiol action of the alkoxyphenylethylamines. An examn. of 4 other possible methods of prepn. indicated that the only feasible one was that based on the scheme $\text{RCHO} (+ \text{CH}_3(\text{CO}_2\text{H})_2 \rightarrow \text{RCH}_2\text{CHCO}_2\text{H} (+ \text{H}_2) \rightarrow \text{RCH}_2\text{CH}_2\text{CO}_2\text{H} (+ \text{SOCl}_2) \rightarrow \text{RCH}_2\text{CH}_2\text{COCl} (+ \text{NH}_3) \rightarrow \text{RCH}_2\text{CH}_2\text{CONH}_2 (+ \text{KOH}) \rightarrow \text{RCH}_2\text{CH}_2\text{NH}_2$. The m and p -mono- and the 2,3-, 3,4- and 3,5-di-MeO compds have already been prepd by this method but in all attempts to prep I it had hitherto been impossible to carry out the last step. S and II have now found, however, that under certain conditions not only I but also its isomers can readily be obtained in very satisfactory yield by using concd. NaOBr for the Hofmann degradation. The condensation of the aldehydes with $\text{CH}_3(\text{CO}_2\text{H})_2$ was effected in 2–3 parts C₆H₅N (distd over KOH) and 1–5 cc piperidine per mol of aldehyde, com. $\text{CH}_3(\text{CO}_2\text{H})_2$ can be used directly and a 20% excess is quite sufficient. The splitting off of CO₂ from the primary di-CO₂H acid is almost quant. on the H₂O bath and only in a few cases was it necessary to boil the C₆H₅N soln a short while. The yield of $\text{RCH}_2\text{CHCO}_2\text{H}$ was in general above 80%. The reduction of the unsatd. acid was in all cases easily carried out with an excess of Na-Hg, usually on the H₂O bath, a higher temp was required only where the unsatd. acid was difficultly sol in NaOH, and a tanning down of the alkyl with AcOH was necessary only with 3,4,5-(MeO)₃C₆H₂CH₂CHCO₂H. The chlorides $\text{RCH}_2\text{CH}_2\text{COCl}$ were obtained by heating the acids in CHCl_3 4–5 hrs on the H₂O bath with twice the calcd. amt of SOCl_2 ; they were not isolated but the reaction mixts were concd to about 0.5 vol *in vacuo* and added, with cooling, to concd. NH₄OH contg. about 10% NaOH (only 2.1 (MeO)₃C₆H₂CH₂CH₂CONH₂ could not be obtained in this way). For the successful degradation of the amides to the amines, it is necessary to have a very pure amide. The statement in the literature that NaOCl gives better yields than NaOBr was confirmed for only some of the amides (the best results were obtained with a soln prepd by passing 5.5 g Cl into 10 g. NaOH in 100 cc H₂O). The behavior of the amides in the Hofmann degradation depends greatly on the position and nature of the nucleus substituents. The products are best worked up by distn. *in vacuo*. The physiol action of the amines prepd will be described later by Hesse and Lange. *o*-MeOC₆H₃(CH₂)₂CHCO₂H, m 183° (80% from *o*-MeOC₆H₃CHO), gave 99% MeOC₆H₃(CH₂)₂CH₂CH₂OH, m 102° (literature, 85–6°), the amide, m 111° (86.3% yield), gave with NaOCl 35.2% *o*-MeOC₆H₃(CH₂)₂CH₂NH₂, b_p 115–20° in the vacuum of a H₂O pump (HCl salt, m 113°, sulfate, m 230°). *m*-HOC₆H₄CHO, m 100° (69% by diazotization of *m*-H₂NCC₆H₄CHO, prepd from the NO₂ compd with FeSO_4), yielded 80% of the MeOC₆H₄CHO which gave 69% MeOC₆H₄(CH₂)₂CHCO₂H, m 177°, the satd. acid, m 50° (80% yield) gave 27.5% of the amide, b_p 222–5°, m 56°, which yielded with NaOCl 61% (with NaOCl 10%); of *m*-MeOC₆H₄(CH₂)₂CH₂NH₂, b_p 124°. *p*-MeOC₆H₄(CH₂)₂CH₂NH₂, HCl , m 210°, was obtained in 43% yield with NaOCl from the amide which itself was obtained in 75% yield through the MeOC₆H₄(CH₂)₂CHCO₂H (92.5%) and the satd. acid (81%). *p*-iso-PrC₆H₄(CH₂)₂CHCO₂H, m 165° (81%); satd. aekl, m 73° (98%); amide, m 142°.

(81.5%), boils explosively with NaOCl at 80° and at 85° gives only a little β (*p*-isopropylphenyl)ethylamine, b. 160° in the vacuum of a H₂O pump (HCl salt, m 270°). Vanillin Et ether, m 63° (85% from vanillin and alk. FeSO₄), gave 80% 3-methoxy-4-ethoxycinnamic acid, m 205°; propionic acid, m 130° (79%), amide, m 124° (33.9%), gave with NaOCl the ethylamine, b. 165° (HCl salt, m 120°) m C₁₁H₁₃(OMe)₂, b. 117-8° (90% from m C₁₁H₁₃(OH)₂ with alk. Me₂SO₄), gave in PtO with HCN, HCl and ZnCl₂ 76% of 2,4 (MeO)₂C₆H₃CHO; 2,4-(MeO)₂C₆H₃CH₂CH₂CO₂H, m 138° (90%), propionic acid, m 105° (91%); attempts to prep the amide through the chloride gave a red brown oil which decomposed on distn. in vacuo 3,4 C₁₁H₁₃O₂C₆H₃CH₂CH₂CO₂H, m 232° (94%), propionic acid (92%); amide (83.5%), gave with NaOCl 37% of the ethylamine-HCl, m 208°, 1,2,3-C₁₁H₁₃(OMe)₃, b. 140° (65% with alk. Me₂SO₄), gives with HCN and AlCl₃ in C₆H₆ 53% 2,3,4 (MeO)₃C₆H₃CHO, b. 170°. 2,3,4-Trimethoxycinnamic acid, m 172° (94%), propionic acid, m. 78° (65%), amide, m 171° (63.5%), gives with NaOBr 71.5% β (2,3,4 trimethoxyphenyl)-ethylamine, b. 167°, quickly forms the carbonate in the air. 3,4,5-(MeO)₃C₆H₃CO₂H (86% from (HO)₃C₆H₃CO₂H with alk. Me₂SO₄), chloride (80% with PCl₅) in boiling xylene (distd from Na) with Pd-BaSO₄ and H gave in 30 hrs 80% 3,4,5-(MeO)₃C₆H₃CH₂CH₂CO₂H, m 74°. cinnamic acid, m 124° (80%), β (3,4,5 trimethoxyphenyl)propionic acid, m 98° (88% yield), amide, m 103° (60%), gave with NaOBr (52 g Br and 150 g NaOH in 375 cc H₂O) 66% I, b. 180° (HCl salt, m 181°). β BrC₆H₃(OMe)₃, b. 120° (82.9% from PhOMe in AcOH with Br vapors) gave in Et₂O with Mg (activated with I) and ethylene oxide and subsequent decompos with HCl 31.1% of *p*-MeOC₆H₄CH₂CH₂OH, b. 145-60° in the vacuum of the H₂O pump, and this with boiling 47.5% HBr yielded 45% of the bromide, b. 140-60°, from which, when refluxed with a mixt. of C₆H₅(CO)₂NH and 0.5 mol K₂CO₃ until no liquid condensed in the condenser tube, then treated in alc with N₂H₄, H₂O and finally with HCl gas, was obtained 11.8% *p*-MeOC₆H₄CH₂CH₂NH₂, HCl, m 211°. 4-Bromoveratrole, *p*-BrC₆H₄OH, *p*-BrC₆H₄OAc, *o*- and *p*-BrC₆H₄NO₂ and *o* C₆H₄Cl₂ could not be made to react with Mg in either Et₂O, (iso-Am)₂O or otherwise

o-Toluic acid. H T. CLARKE AND E R TAYLOR. *Org Syntheses* XI, 96-7 (1931) — Sapon of *o*-MeC₆H₄CN with 75% H₂SO₄ gives 80-90% of *o*-MeC₆H₄CO₂H, m 103-3°; *p*-MeC₆H₄CO₂H is obtained in the same yield by the same method. C J WEST

Preparation of aromatic acid amides. C H KAO AND SHAO-YUAN MA. *J Chem Soc* 1931, 443-4, cf C A 25, 1220 — If 100 g aromatic acid and 150 cc 90% AcOH are used with 90 g (NH₄)₂CO₃ the following yields of amides are obtained: *p*-O₂NC₆H₄CO₂H 60 g, *m* NO₂ deriv, 58 g, *o*-NO₂ deriv, 15 g, *p* Cl deriv, 79 g, benzoic acid, 77.5 g (71 g with 60 g (NH₄)₂CO₃), C₆H₅(CO)₂O (60 g), 90 g imide; *o* HOC₆H₄CO₂H, 0, PhCH₂CO₂H and PhCH₂CH₂CO₂H give the amides quantitatively, which are in part dehydrated to the nitriles, PhCH₂CHCO₂H (100 g) gives 20 g amide (C₆H₅CO₂H); (100 g) gives 71 g of the imide

Nitration of 6-methoxy-*m*-toluic acid. A correction. JOHN L SIMONSEN. *J Chem Soc* 1931, 444 — Nitration of 3,6 Me(MeO)₂C₆H₃CO₂H gives the 6-NO₂ deriv. and 6-nitro-*o*-tolyl Me ether, m 69-70°, and not the 6-NO₂ deriv., as reported in C A 13, 127.

C. J. WEST
A method for the preparation of phenyl-, *o*-chlorophenyl- and *p*-chlorophenylmalonic acids. D IVANOFF AND A SPASSOFF. *Bull soc chim* [4], 49, 19-23 (1931) — Grignard syntheses designed to produce ketones frequently give hydrocarbons and a complex (I) instead. I and S made use of I in synthesizing phenyl-, (II), *o*-chlorophenyl-, (III), and *p*-chlorophenyl malonic acids (IV) according to the scheme: PhCH₂CO₂MgX + RMgX → PhCH(RMgX)CO₂MgX (I) + RH, I + CO₂ → PhCH(CO₂MgX)₂, which on hydrolysis yields II. I and S found that all aliphatic, arylaliphatic and hydroaromatic, but only certain aromatic organomagnesium compounds reacted in the above manner, and they have studied the reaction quantitatively to find out which gives the best yields. With PhCH₂CO₂MgCl the yields of mixed acids (in g) and of II (in %) from the following Grignard reagents were, resp.: EtMgBr, 15.25, 62.5, PrMgCl, 11.90, 45, iso-PrMgCl, 14.10, 65.6, iso-PrMgBr, 10.90, 48.9; BuMgBr, 11.00, 42.2, C₆H₅MgBr, 9.8, 40, *o*-MeC₆H₄MgBr, 11.80, 50.6, *o*-C₆H₄MgBr, 13.00, 53.3, PhCH₂MgCl, 12.40, 31. MeMgI gave a little gas evolution but hardly a trace of II. The procedure in all cases was to treat 0.1 g mol. of the Grignard reagent (e.g., PhCH₂CO₂MgCl) in 100 cc ether at 0° with dry CO₂ in excess in a closed system, expel the excess CO₂ by evap. half the ether, add 0.15 g mol. of the 2nd Grignard reagent (e.g., EtMgBr), reflux on the water bath, treat again with dry CO₂ in excess at 0°, hydrolyze, ext. the org. acids with ether, ext. the org. acids from the ether with dil. alkali, acidify, ext. again with ether, dry and weigh. This product contains

the corresponding phenylmalonic acid phenylacetic acid and the acid from the 2nd Grignard reagent. A weighed sample heated to 150–60° in a flask gave CO₂ and the corresponding phenylacetic acid. The wt. of CO₂ collected gave the % of the corresponding phenylmalonic acid. Only the phenylmalonic acids are sol. in benzene and can be sepd. easily in the pure state by extr. with benzene. With *o*-ClC₆H₄CH₂CO₂MgCl, *iso*-PrMgCl gave 10.5 g. mixed acids and 46.2% III; PhMgBr gave 10.5 (?) g. mixed acids and 52.8% of III. III m. 139° (decompn.), heating quantitatively produces CO₂ and *o*-ClC₆H₄CH₂CO₂H. With *p*-ClC₆H₄CH₂CO₂MgCl *iso*-PrMgCl gave 13.5 g. mixed acids and 56.4% of IV; *o*-MeC₆H₄MgBr gave 13.0 g. mixed acids and 48.3 g. IV. IV m. 163° (decompn.) and on heating yields quantitatively CO₂ and *p*-ClC₆H₄CH₂CO₂H. The hydrocarbons evolved in the prepn. of I by the use of EtMgBr, PrMgBr, and *iso*-PrMgBr were collected, washed free of ether with concd. H₂SO₄ and burned with excess O for quant. detn. The amts. found checked very well with those calcd.

ODEN E. SHEPPARD

Dibenzylsuccinic acids. P. CORDIER *Compt. rend.* 192, 361–3 (1931).—Stobbe (*Ber.* 27, 2406 (1894); 37, 2241, 2668 (1904), *C. A.* 3, 438, 651) prepd. dibenzylsuccinic acid, [CH(CH₂Ph)CO₂]₂, and reported obtaining 2 isomeric acids: the one (m. 203°) (I) he considered inactive by internal compensation and the other (m. 204°) as racemic. Further, S. indicated that the inactive compd. gave an anhydride m. 104° (II) which under the action of heat yielded an anhydride m. 128° (III), while the *dl*-compd. gave an anhydride m. 155° which under the action of heat yielded an anhydride m. 140°. Following the same procedure, C. obtained I, which gave 2 anhydrides, the first (II) being obtained by moderate treatment with Ac₂O, and the second (m. 125°) (IV) by more vigorous treatment of II. His 2nd anhydride (IV) seemed to correspond to the III of S. However, it did not give the original dibenzylsuccinic acid on hydration but a new isomeric acid (m. 172°) (V) which could easily be converted back to IV. C. found that I was inactive and could not be resolved into active components by strychnine sulfate, V was also inactive but it gave, by the same treatment, 2 active components, the one with α_d = –20.5°, the other, less pure, with α_d 22.8°. These fast 2 acids are much more sol. in water than V, and the m. ps. are much lower (about 130°).

ALBERT L. RAWLINS

Esters of orthophenylacetic acid. PETER P. T. SAI, SHAO YUAM MA and CHUNG HSU KAO, *J. Chem. Soc.* 1931, 305–7.—PhCH₂CN and abs. MeOH (EtOH) in an equal vol. of abs. Et₂O with 1.25 mols. HCl at 0° give 85% of PhCH₂C(OMe)NH₂ HCl (or the Et deriv.), with 2 mols. abs. alc. at room temp. for 2 weeks, these results 40–5% of PhCH₂C(OMe)(OR), (I) or PhCH₂C(OEt)(OR), (II); they are colorless liquids with a pleasant odor. Derivs. of I (R =) (b. p. at atm. pressure, d_4^{25} and n_D^{25} are given): *Me*, 216–8°, 1.0651, 1.5075, Et, 224–6°, 1.0356, 1.5000, *Pr*, 239–42°, 1.0109, 1.4930; *iso*-*Pr*, 227–9°, 1.0079, 1.4913, *Bu*, 254–7°, 0.9053, 1.4911; *iso*-*Bu*, 245–8°, 0.9029, 1.4898; *iso*-*Am*, 260–5°, 0.9880, 1.4900. Derivs. of II: *Me*, 217–9°, 1.0640, 1.5080, Et, 225–7°, 1.0308, 1.4997; *Pr*, 238–41°, 1.0094, 1.4967, *iso*-*Pr*, 228–30°, 1.0030, 1.4908; *Bu*, 254–7°, 0.9974, 1.4916, *iso*-*Bu*, 248–51°, 0.9867, 1.4883; *iso*-*Am*, 260–5°, 0.9867, 1.4887.

C. J. WEST

Ethyl phenylcyanopyruvate. ROGER ADAMS and H. O. CALVERY, *Org. Syntheses* XI, 40–1 (1931).—PhCH₂CN, (CO₂Et)₂, and EtONa, followed by acidification, give 69–75% of PhCH(CN)COCO₂Et, m. 130°.

C. J. WEST

New method for the preparation of β -ketonic esters. (MILLER) S. GRATEAU, *Compt. rend.* 191, 947–9 (1930).—The acid chloride obtained by treating EtOCO(CH₂)₂CO₂H with SOCl₂ reacts with PhI in the presence of AlCl₃ to give PhCO(CH₂)₂CO₂Et (I) in 80% yield. I, b. p. 164°, is sapon. to give PhCO(CH₂)₂CO₂H, m. 71°, which forms a semicarbazone m. 183°. I treated with EtONa gives CH₃(CH₂)₂CO CH-

COPh, b. p. 137° (semicarbazone, m. 223°, Maquenne block). Reduction of I by the Clemmensen method gives Ph(CH₂)₂CO₂Et (II), b. p. 162–4°. Sapon. of II gives Ph(CH₂)₂CO₂H, b. p. 180–8°, m. 11°, anilide, m. 80°, *p*-toluide, m. 78°. C. H. PRET

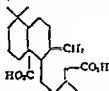
Lichen substances. VI. Gyrophoric acid. YASUHIKO ASAHINA and MASATOSHI WATANABE, *Ber.* 63B, 3044–8 (1930), cf. *C. A.* 23, 4477.—The conclusion that gyrophoric acid (I) is a tripeptide, C₁₅H₂₁O₁₀, of orsellinic acid is confirmed by the present investigation. The m. p. of I from various sources (*Gyrophora esculenta* Miyoshi, *Lobaria pulmonaria* Hofm.) can be raised by repeated crystn. to 220° (202–3° has hitherto generally been given as the m. p.). Titration of the I.H₂O in Me₂CO with bromothymol blue as indicator gave a mol. wt. of 477.5–7.8 (a parallel titration with lecanonic acid gave 340.1). *Tetra-Ac* deriv. of I, from I, Ac₂O and a drop of concd. H₂SO₄, m. 228°, gives no color in alc. with FeCl₃ or bleaching powder soln., mol. wt.

(by titration) 636-641.1, in camphor (Rast) 617-32. *Tetra(chloroacetyl) deriv.*, from I in $\text{C}_6\text{H}_5\text{N}$ and $\text{ClCH}_2\text{CO}_2\text{H}$ treated alternately with $\text{C}_6\text{H}_5\text{N}$ and $\text{ClCH}_2\text{CO}_2\text{H}$, cryst. powder with 2 H_2O , m. 163-4°, gives no color in aq. Me_2CO with FeCl_3 or bleaching powder, mol. wt. (anhyd.) in boiling Me_2CO 704-815. *Quinine salt* of I, faintly yellowish cryst. powder with 1 H_2O , m. 162°, mol. wt. (anhyd.) in boiling Me_2CO 766.2-789.7. *Me tetramethylgyrophorate* (0.6 g from 1 g I in Et_2O with CH_3N_3), m. 156-7° (previously given as 157°), insol in Na_2CO_3 , gives no color in alc. with FeCl_3 or bleaching powder, mol. wt. in freezing EtOH 510.4. Since, according to Heise, umbelliferic acid, $\text{C}_{11}\text{H}_{10}\text{O}_5/\text{O}(\text{Me})$ (II), which is also found with I in many *Gyrophorae*, is decompd. by $\text{Ba}(\text{OH})_2$ into CO_2 , o-cresol and umbelliferonic acid, isomeric with evernic acid, the relation of I to II seems to be the same as that of lecanoric acid to evernic

acid, viz., $\text{Me}(\text{HO})(\text{XO})\text{C}_6\text{H}_4\text{CO}_2\text{OC}_6\text{H}_4\text{Me}(\text{OH})\text{CO}_2\text{OC}_6\text{H}_4\text{Me}(\text{OH})\text{CO}_2\text{H}$, where $\text{X} = \text{H}$ for I and Me for II. Also in *J Pharm Soc Japan* 51, 104-9 (1931) C. A. R.

Symmetrical and unsymmetrical *o*-phthalyl chlorides. EAWY ORY *Org Syntheses* XI, 85-91 (1931)—*o*- $\text{C}_6\text{H}_4(\text{COCl})_2$ results from $\text{C}_6\text{H}_4(\text{CO})_2\text{O}$ and PCl_5 in 92% yield. AlCl_3 gives 72% of the unsym. chloride C. J. WEST

Higher terpene compounds. XLIII. The Bouveault reduction of the esters of agathic- and isogathedicarboxylic acids. Transformation of the latter into a new methylpimarane. L. RUTICKA AND J. R. HOSKING *Helv Chim Acta* 14, 203-19 (1931), cf. C. A. 25, 1232—R. and H. suggest I and II as the skeleton formulas for agathic- (III) and isogathedicarboxylic acids (IV). The facts upon which these formu-



(I)



(II)

las are based are as follows: (a) I yields 1,5,6- $\text{C}_{11}\text{H}_{16}\text{Me}_3$ (V) and 1,7-dimethylpimarane (VI) upon reduction with Se . II also yields VI but no V. (b) The 2 CO_2H groups in III are different—one of the CO_2Me groups in the di-Me ester of III is easily hydrolyzed, while the other resists hydrolysis. The esters of IV, isogathic acid (VII), and isogathic acid also resist hydrolysis. (c) Upon reduction of III with Na in alc. at 115-20° a mixt. was obtained of which 20% was neutral and 70% acidic. Of the neutral portion 25% consisted of *mono-Me dihydroxyphthalate*, $\text{C}_{11}\text{H}_{14}(\text{CO}_2\text{Me})\text{CH}_2\text{OH}$, b.p. 123-5°, d_4^{25} 1.029, n_D^{25} 1.556, M_D 97.57, (calcd. 97.47). The remaining 2% contained *dihydroxydihydroagathene*, $\text{C}_{11}\text{H}_{14}(\text{CH}_2\text{OH})_2$, m. 112-3°, which absorbed 1 mol. of H_2 in the presence of PtO to give the satd. compd. *dihydroxytetrahydroagathene*, $\text{C}_{11}\text{H}_{16}(\text{CH}_2\text{OH})_2$, m. 107-8°. The acidic portion was distd. in *vacuo* and yielded the Me ester of VII; upon repeated redistn. of the acidic portion of the distillate *mono-Me dihydrogathedicarboxylic acid*, $\text{C}_{11}\text{H}_{14}(\text{CO}_2\text{Me})\text{CO}_2\text{H}$, b.p. 210-2°, was obtained. The stability of this compd. in comparison to that of its precursor, $\text{C}_{11}\text{H}_{14}(\text{CO}_2\text{Me})_2$ (VIII), illustrates the effect of the double bond in close proximity to the CO_2H group. VIII contains a double bond which is easily reduced, close to a CO_2Me group which is also easily reduced and hydrolyzed. This indicates that the C atoms on either side of the double bond are unsubstituted. Reduction of the di-Me ester of IV with Na and alc. yields *Me dihydroxyphthalate*, $\text{C}_{11}\text{H}_{14}(\text{CO}_2\text{Me})\text{CH}_2\text{OH}$ (IX), m. 125-6°, and *dihydroxyagathene*, m. 172-3°, in almost equal amts. A small quantity of a dehydro deriv. of IX, $\text{C}_{11}\text{H}_{12}\text{O}_2$, b.p. 155-60°, was also isolated. Treatment of IX with 4 vols. of 90% HCO_2H yields a mixt. (b.p. 155-8°) of the esters $\text{C}_{11}\text{H}_{14}(\text{CO}_2\text{Me})_2$ and $\text{C}_{11}\text{H}_{14}(\text{CO}_2\text{Me})\text{H}$. This mixt. when reduced with Se yielded $\text{C}_{11}\text{H}_{16}$ (X), m. 142-3° (picrate m. 161-3°). Oxidation of X with CrO_3 gave the *quinone*, m. 134°, from which the *quinonolone compd.*, m. 131-2°, was obtained. It is believed that X is a homolog of VI. XLIV. The constitution of pimarane. L. RUTICKA, G. B. R. DE GRAAF AND J. R. HOSKING *Ibid* 223-9—Pimarane (XI) and retene (XII) were each oxidized in parallel expts. with $\text{K}_2\text{Fe}(\text{CN})_6$ and KOH , and the mixt. of resulting acids, sepd. by means of cold MeOH , was esterified with CH_3N_3 . In both cases *tri-Me 1,2,1'-3'-dihydroxyphthalate* (XIII), m. 153-4°, and *1,7-pimarane dicarboxylic acid*, m. 151-2°, were formed. Likewise the *o*-quinones of XI and XII were oxidized with KMnO_4 . The XII quinone yielded *3'-hydroxyisopropyl-1,2,1'*-

biphenyltricarboxylic acid (XIV), m 188° (Bucher, *C. A.* 4, 1465, obtained the same compd in a similar fashion but he gave no m p), and an amorphous substance which gave upon further oxidation with HNO_3 (d 15) and esterification of the product, 1,2,3- (XV) and 1,2,4 $\text{C}_6\text{H}_3(\text{CO}_2\text{Me})_3$. The XI quinone and XIV upon oxidation yielded XIII and XV. From these facts it is concluded that XI is 1,7-dimethylphenanthrene.

I. M. LEVINE

Catalytic influences in three-carbon tautomerism. II. The action of acid catalysts on ketones. GEORGE A. R. KAY, *J. Chem. Soc.* 1931, 248-51, cf. *C. A.* 23, 4454. The action of H_2SO_4 , HCl and $(\text{CO}_2\text{H})_2$ on cyclohexylidene- (I) and cyclohexenylacetone and on isopulegone is reported. The acids are less rapid in their action than the catalysts previously employed. e. g., N aq. EtOH H_2SO_4 at 25° produces only 22% change in I in 4 hrs, which is equiv to about 4-min treatment with N EtONa . Longer treatment gives rise to extensive side reactions (addn), at 100°, however, isomerization proceeds smoothly and side reactions are negligible in the time required to reach equilibrium. Heating for 0.5 hr. is sufficient to produce the same equil. from either ketone by means of N H_2O EtOH H_2SO_4 . With HCl the same equil. p is reached but side reactions occur more readily and a Cl-contg. high boiling fraction is obtained. $(\text{CO}_2\text{H})_2$ is much slower in action. The new value for the equil. p is 23% of α,β -ketone. Isopulegone is rapidly isomerized to pulegone by hot aq. H_2SO_4 , the equil. p being coincident, within exptl. error, with pure pulegone. Although pulegone is readily isolated from the equil. mixt. as the semicarbazone, its phys. properties differ somewhat from those of pure pulegone and it may be that the equil. of isopulegone is a less simple process than hitherto supposed. Results of 18 expts. are given in tabular form. C. J. W.

The energy value of the ozonide linkage determined from the ozonide of α -terpineol. E. BRINER, M. MOTTIER AND H. PAILLARD, *Helv. Chim. Acta* 13, 1030-5 (1930).—Comparison of the heats of combustion of a compd. and its ozonide would give the energy equiv. of the ozonide linkage. The heat of formation would appear to be strongly exothermic since the ozonization of C_7H_{14} , etc., is accompanied by a considerable rise in temp. but this rise may be due to decompn. of the ozonides as well as to the heat of formation of the ozonide. For the detn., the ozonide of α -terpineol (I) was chosen because it is stable and a solid. I was ozonized in specially dried hexane (II), the II was then decanted off and the I ozonide was placed in a vacuum desiccator to remove the residual II. Small quantities of an unidentified gas were continuously given off and it was found that the heat of combustion decreased as the age of the ozonide increased. The av. heat of combustion of freshly prepd. I ozonide was 1445 cal. per g. mol. and that of I was 1501 cal., from which the heat of ozonization of I was found to be 90 cal. This value is probably significant though not rigorously exact. Since it indicates a lower energy content than that of the components which enter into the ozonides, there would appear to be no advantage in the suggested employment of ozonides as explosives unless ozonization alters the mol. to make disruption of its linkages easier. C. H. PEET

Oil of Eucalyptus dives. K. STEPHAN AND MARG. DCKER, *J. prakt. Chem.* 129, 145-50 (1931).—The oil from *Eucalyptus dives* contains 35% of pipentone, the oil also contains about 4.5% of terpinen-4-ol and also an acid, $\text{C}_{18}\text{H}_{34}\text{O}_4$, m 109-12°. Oxidation of pipentone with FeCl_3 gives 90% of thymol, oxidation with neutral KMnO_4 gives diosphenol, but the yield is too small to make this a technical method of prepn. C. J. WEST

Thiobenzophenone. H. STAUDINGER AND H. FREUDENBERGER, *Org. Syntheses* XI, 94 5 (1931).— Ph_2CCl_2 and NaSH in EtOH give 42-9% of Ph_2CS , m 53-4°; if the chloride is added to the NaSH a 70% yield of the disulfide is obtained. C. J. WEST

4-Nitro-4'-methoxystilbene. A correction. J. T. HEWITT AND WM. LEWCOCK, *J. Chem. Soc.* 1931, 444.—4-Nitro-4'-methoxystilbene m 132°, not 162° as reported in *C. A.* 6, 3264. C. J. WEST

α -Chloronaphthalene. V. The chlorination of naphthalene in solution. P. FERRERO AND J. CORBAZ, *Helv. Chim. Acta* 13, 1009-25 (1930), cf. *C. A.* 22, 2938; 23, 4944; 24, 106, 1652.—The study of the chlorination of C_{10}H_8 (I) was continued in a variety of solvents to establish, if possible, a series which would either be protected by the chlorination product or which would prevent the chlorination of I. In every case, 2 mols. of I was dissolved in 500 g. of solvent and Cl_2 passed in at the rate of 35 g. per hr. When a catalyst was used, a wt. was taken equal to 0.5% of the I present. Below the limits at which partial oxidation occurs, higher temps. favor monohalogenation. I_2 always favors monohalogenation while FeCl_3 favors polyhalogenation. The concn. mentioned above was about the optimum, less solvent approaching the condition of no solvent with consequent polyhalogenation and more solvent resulting in gradual

chlorination of the solvent and polyhalogenation of I. The theoretical amt. of Cl_2 should be added but the rate of addn. is immaterial so long as the Cl_2 is thoroughly dispersed throughout the reaction mixt. The chem. nature of the solvent is more important than the amt. present or the phys. properties of the soln. Solvents which cannot be acted upon by Cl_2 (e. g., CCl_4) favor a considerable formation of poly Cl derivs. and are less desirable than no solvent. Solvents which might be chlorinated (a) (e. g., PhH or PhCl) may be protected from chlorination by I even in the presence of I_2 and are the most satisfactory solvents for use in chlorination or (b) (e. g., PhMe or C_6H_{11}) they may protect the I against chlorination and hence be entirely unsatisfactory solvents. PhH , PhCl and AcOH are the most satisfactory solvents investigated. PhMe , $\alpha\text{-C}_6\text{H}_4\text{Cl}$, PhNO_2 , $\text{C}_6\text{H}_5\text{Cl}$, $\text{C}_6\text{H}_5\text{Br}$, EtOH , AcOEt and Li_2O are all unsatisfactory for reasons mentioned above.

C. H. PEET

Naphthalene series. III. Preparation of 3-amino-2-naphthyl methyl ether. GEV BAMSISIAH JAMBHERWALLA, SIDNEY HOLT AND FARDEICK A. MASON. *J. Chem. Soc.* 1931, 373-7. $\text{C}_6\text{H}_5\text{Cl}$ 19, 437—2,3- $\text{HOC}_6\text{H}_4\text{CO}_2\text{H}$ (48 g.) in 100 cc. 20% NaOH with 64 g. Me_2SO , at 15° gives 49.6 g. of 2-methoxy-3-naphthoic acid, pale yellow, m. $133-5^\circ$, the use of too great excess of alkali decreases the yield, the chloride is a transparent yellow gum, m. about 30° (decomps.), which gives with $(\text{NH}_4)_2\text{CO}_3$ in $\text{C}_6\text{H}_5\text{Cl}$, 77% of 2-methoxy-3-naphthamide, m. $172-3^\circ$, 15 g. of the amide in EtOH , treated with KOH , gives not over 3 g. of 3,2- $\text{H}_2\text{NC}_6\text{H}_4\text{OMe}$, m. $109-10^\circ$. The diazo compd. with CuCl gives 82% of 3-chloro-2-naphthyl Me ether, m. 78.5° , HCl gives 3-chloro-2-naphthol, m. $67-68^\circ$, gives a green color with FeCl_3 , 2,3- $\text{MeOC}_6\text{H}_4\text{CO}_2\text{Et}$, b.p. 198° , gives 98% of 2-methoxy-3-naphthylhydrazide (I), m. $121-4^\circ$, it melted slowly when heated on the H_2O bath and then m. $137-8^\circ$, with 2,3- $\text{MeOC}_6\text{H}_4\text{COCl}$ there results 80% of di (2-methoxy-3-naphthyl)hydrazide, m. $219-70^\circ$. I and BaCl_2 give the benzylidene deriv., m. 222.5° (cor.), m. nitrobenzylidene deriv., canary yellow, m. $223-4^\circ$ (cor.), α -isomer, canary yellow, m. $195-6^\circ$ (cor.), I and vanillin give 4'-hydroxy-2,2'-dimethoxybenzylidene-3-naphthylhydrazide, m. $211-2^\circ$ (cor.) I in EtOH HCl with EtNO_2 gives 72% of di 2-methoxy-3-naphthylcarbamate m. $104-5^\circ$ (cor.), hydrolysis with concd. HCl at 120° for 3 hrs. gives quant. 3,2- $\text{H}_2\text{NC}_6\text{H}_4\text{OH}$, 10% H_2SO_4 gives the naphthol and some Me ether, heating with slaked lime for 3 hrs. at 200° gives 85% of the Me ether. 40% NaOH on refluxing 1 hr. gives 92.5% of the Me ether. IV. Preparation and properties of 2-naphthol-3-sulfonic acid. S. HOLT AND P. A. MASON. *Ibid.* 377-81—3,2- $\text{H}_2\text{NC}_6\text{H}_4\text{OMe}$ (14 g.) was diazotized in a mixt. of 48 cc. concd. H_2SO_4 and 160 cc. H_2O at 0° with 2 N NaNO_2 , SO_2 was passed into the diazo soln. until frothing became noticeable and 10 cc. liquid SO_2 were then run in, chem. prepd. Cu powder (30 g.) was then added, a slow stream of SO_2 being maintained, the ppt. was then extd. with 30% K_2CO_3 and the ext. acidified with concd. HCl , giving nearly quant. 2-methoxynaphthalene-3-sulfonic acid (I), m. $173-4^\circ$. Oxidation with KMnO_4 in Me_2CO gives nearly quant. the K salt (II) of the SO_3H acid, oxidation in H_2O gives 75% of a less pure product. The alk. soln. of the salt gives no color with 2,4- $(\text{O}_2\text{N})_2\text{-C}_6\text{H}_3\text{N}_2\text{Cl}_2$. 2-Methoxynaphthalene-3-sulfonyl chloride m. $137-8^\circ$, the amide m. 113° , the anilide m. $173-4^\circ$. I and 2,4- $(\text{O}_2\text{N})_2\text{C}_6\text{H}_3\text{N}_2\text{Cl}_2$ in EtOH-AcONa give 2,4-dinitrophenyl 2-methoxy-3-naphthyl sulfone m. $213-4^\circ$. II and HCl (1:1), refluxed 1 hr., give 93 of 2-naphthol-3-sulfonic acid (III), which seps with 1 mol. H_2O ; the Na salt also seps. with 1 mol. H_2O , the chloride is amorphous, the PhNH_2 salt m. $241-2^\circ$; the $\alpha\text{-C}_6\text{H}_4\text{NH}_2$ salt m. $247-8^\circ$, the sulfonamide, yellow, m. 110° ; 2-naphthol-3-sulfonamide, buff., m. 112° (decomps.) 1-Nitroso-2-naphthol-3-sulfonic acid, yellow, decomps. 268° , FeCl_3 gives a deep brown color, it dyes Fe-mordanted cotton a tobacco-brown. III yields azo dyes as follows: PhNH_2 , red needles with blue reflex; $\beta\text{-O}_2\text{NC}_6\text{H}_4\text{NH}_2$, red needles, m-xylidine, red powder; $\alpha\text{-C}_6\text{H}_4\text{NH}_2$, dianiline. Dyeings with these on wool were deeper in all cases than the corresponding shades obtained from Schaefer acid. The peaks of the absorption curves of these dyes are given.

C. J. WEST

Alkamine and ether bases of the tetralin and hydrundene series. JULIUS V. BRAUN AND KARL WEISSBACH. *Ber.* 63B, 3052-9 (1930)—It was recently shown (*C. A.* 25, 1812) with $\text{CH}_3\text{O}_2\text{C}_6\text{H}_4\text{CH}(\text{OH})\text{CHBrMe}$ (I) that it is possible to det., by methylation with Me_2SO , whether in the alkamines formed (through the α -oxide) by treating I with an amine RNH_2 , the RNH is in the β position originally occupied by the Br and the OH in the α position, as in the I, or whether, when the intermediate α -oxide ring is opened, the NHR group goes to the α -position and the HO to the β -position. This method has now been used to answer the long unanswered question as to the structure of the alkamines (II) obtained from α,β -dibromotetralin through the bromohydroxy compd. with various amines. Methylation of the alkamine obtained

with NHMe_2 and treatment with KI gives a *quaternary iodide* (III) different from that (IV) obtained by replacing the Br in α -methoxy- β -bromotetralin by NMe_2 and adding MeI . While IV decomps into a tertiary base and MeI on heating, III gives NMe_2 , HI and the *Me ether* (V), $\text{C}_{10}\text{H}_7\text{CH}=\text{C}(\text{OMe})\text{CH}_2\text{CH}_2$, of the enol form of the tetralone;

V readily takes up 2 atoms H and on dehydrogenation with Br yields β - $\text{C}_{10}\text{H}_7\text{OMe}$, showing that in the II the HO occupies the β position. Analogous expts in the indene series showed that, contrary to the hitherto generally accepted view, the alkamines (VI) obtained from α -hydroxy β -bromohydrindene have the HO group in the β position. Partly to confirm these results by a different method, and partly to obtain the α - HO isomers of the I and VI for pharmacol. comparison with the I, attempts were made to prep these α - HO isomers by demethylation of the corresponding α - MeO compds. These attempts were successful only with α -methoxy- β -methylaminotetralin, the resulting alkamine (VII) being different from the II from α -hydroxy- β -bromotetralin with MeNH_2 . When the basic substituent of the α - MeO compds. is NMe_2 or NC_4H_9 , the only definite products of demethylation in the tetralin series are the unsatd bases $\text{C}_{10}\text{H}_7\text{CH}=\text{C}(\text{NR}_2)\text{CH}_2\text{CH}_2$ (VIII), and the analogous unsatd bases are the chief

products in the indene series even where the basic substituent is NHMe . VII, which may be considered as a cyclic ephedrine, acts pharmacologically like ephedrine, although not as powerfully. III, m 168°. V, b_{11} 136°, d_4^{25} 1.0674, turns yellow and thickens on standing several weeks, gives the same blue color as β -tetralone when shaken in alk. soln in the air, gives the tetralone, m . 132–3°, when shaken several hrs with 2 N HCl , is reduced by H and Pd to α - β -tetralyl *Me ether*, b_{11} 123–4°. α -Methoxy- β -dimethylaminotetralin, from the β - Br compd. and NHMe_2 in C_6H_6 at 100°, yellow, b_{11} 147–9° (30% yield), HCl salt, m 184°, *picrate*, m 192°; *methiodide* (IV), m 210°, depresses the m p of III to 135–48°. α -Methoxy- β -methylaminotetralin (50% yield), b_{11} 144–6°; HCl salt, m 228°, *picrate*, m 150°. Exhaustive methylation of the base gives IV. α -Methoxy- β -piperidinotetralin, b_{11} 191–3°, HCl salt, m 230°, *picrate*, m 153°. α -Hydroxy- β -methylaminotetralin (VII) (60% from the α - MeO compd with 66% HBr at 55–60°), m . 77–9° (HCl salt, m 206°), if the heating with HBr is carried out at 100°, the product is the unsatd. base, b_{11} 162–4°, of the type VIII; it forms a hygroscopic HCl salt, m . 148–50°, and a yellow *picrate* m 149°. The unsatd dimethylamino base (VIII), b_{11} 102–4°, is obtained even when the temp in the heating with HBr is not raised above 40–5°. It darkens and thickens in the air and forms an oily HCl salt and a *picrate* m 148°. The piperidino analog, b_{11} 168–70°, m . 40°, is likewise formed at 40–5°, *picrate*, yellow, m 149°. β -Hydroxy α -methylaminohydrindene (VI), m . 130°, HCl salt, m 175°, *picrate*, m 153°. Methylation with Me_2SO_4 and treatment with KI gives a quaternary iodide, m 187°, which splits off NMe_2 , HI only slowly and incompletely above its m p., and gives a light yellow oil, b_{11} 114–6°, having approx. the compn of β -hydrindone enol *Me ether* and giving β -hydrindone when shaken with H_2O . A more complete decompn can be effected by distn. *in vacuo*, which gives β -hydrindone and a base $\text{C}_{10}\text{H}_7\text{ON}$, regenerating the iodide with MeI and different from IX (below), it is therefore probably α -dimethylamino- β -methoxyhydrindene. It b_{11} 132°, *picrate*, m 159°. The β , α -isomer (IX), from β -bromo- α -methoxyhydrindene (b_{11} 134°, obtained in 90% yield from the di Br compd with boiling MeOH) with Me_2NH in C_6H_6 (yield, 20%), b_{11} 128–9°, rapidly reddens, forms a *methiodide* m 176°, which quantitatively regenerates IX and MeI on dry distn. *Methylamino analog* of IX (35%), b_{11} 127–8°, darkens even more rapidly than IX; *picrate*, m 118°, HCl salt, m 173°. Demethylation of the MeO is much more difficult than in the tetralin series, heating with concd. HBr at 100° for even only 1 hr gives a thick non-distillable basic oil, apparently a polymerization product of the unsatd amine, for by heating 1 hr. at 50°, treating with $\text{NaOH}\cdot\text{K}_2\text{CO}_3$ extg with Et_2O and treating in petr. ether with HCl in Et_2O there is obtained a salt m 165–6° having very nearly the compn. $\text{C}_{10}\text{H}_{11}\text{NCl}$, of the HCl addn product of the unsatd amine. The α -hydroxy β -methylaminohydrindene, m . 77–9°, was finally obtained in small yield by allowing the MeO compd. to stand 40 hrs. with fuming HBr at room temp. and then heating 15 min. at 45–50°; its yellow *picrate* m 171°. IX on short heating with HBr gave a basic oil which soon began to thicken and could not be distd.

C. A. R.

1-Amino-2-naphthol hydrochloride. J. B. CONANT AND B. B. CORSON. *Org. Syntheses* XI, 8–11 (1931). —1,2- $\text{C}_{10}\text{H}_7(\text{NO})\text{ONa}$, reduced with $\text{Na}_2\text{S}_2\text{O}_4$ at 60–5°, gives 60–74% 1,2- $\text{C}_{10}\text{H}_7(\text{NH}_2)\text{OH}$.

C. J. WEST

1-Amino-2-naphthol-4-sulfonic acid. LOUIS F. FIESER. *Org. Syntheses* XI, 12–4

(1931) — 1,2- $C_{10}H_7(NO)OH$ in $NaHSO_4$ and $N NaOH$, treated with concd. H_2SO_4 , gives 82-4% of 1,2,4- $H_7N(HIO)C_{10}H_7SO_3H$ C. J. WEST

2-Iodo-3-naphthoic acid. HENRI GOLDSTEIN AND FANEST CORNARUSAI *Helv. Chim. Acta* 14, 399-401 (1931). — Diazotization of 2,3- $C_{10}H_7(NH_2)CO_2H$ and treatment of the product with KI gives 70% of 2-iodo-3-naphthoic acid (I), m. 214°. *Me ester* (85% yield), m. 55°. *Et ester* (87% yield) m. 74°. *amide* (from the action of concd. NH_4OH on the impure chloride prepd. by treatment of I with PCl_5), m. 241° (yield 90%), anilide, m. 235° (yield 60%). I. M. LEVINE

Syntheses of antiseptic derivatives of indan-1,3-dione. I. Interaction of malonyl chloride and of alkylmalonyl chlorides with the methyl ethers of resorcinol and β -naphthol. RAPHAEL BLACK, HERBERT SHAW AND THOMAS K. WALKER. *J. Chem. Soc.* 1931, 272-9 — $CH_3(COCl)_2$ obtained in 36% yield from the acid and PCl_5 with m - $C_6H_4(OMe)_2$ and $AlCl_3$ in $PhNO_2$ gives 20% of 7-hydroxy-5-methoxyindan-1,3-dione, lemon yellow, m. 217°. $MeCH(COCl)_2$, b.p. 45/50°, gives the 2 *Me deriv.*, m. 225-6°, 2 *Et deriv.* pale fawn-colored, m. 192-5°. 2 *Pr deriv.*, pale fawn, m. 190-5°, the Bu deriv. could not be obtained crystalline. β - $C_{10}H_7OMe$ and $CH_3(COCl)_2$ with $AlCl_3$ in $PhNO_2$ give 31% of 5-hydroxy-3-naphthindan-1,3-dione, dull yellow, m. 276-80°, 2 *Et deriv.*, golden brown, m. 215° (20% yield) 2 *Pr deriv.*, bright canary yellow, m. 181°, 2 Bu deriv., yellow, m. 155°. These compounds exert no action on *B. coli communis* when tested by the Rideal-Walker technique and they were also without effect on the proliferation of organisms of the *B. pyocyaneus* group, very definite inhibition of growth was observed with *S. pyogenes* and a no. of Gram positive bacilli. These compounds are to be regarded as possessing selective antiseptic properties distinct from the general disinfectant properties of the phenols from which they are derived. Detailed results are given with *Bacterium C*. The introduction of an alkyl substituent into the 2-position of the dione ring in the parent resorcinol derivative results in a lowering of the antiseptic power, but this depressing effect is counterbalanced to an increasing degree as the length of the substituent alkyl chain is increased. Such lowering effect is absent in the alkylated indandiones derived from β - $C_{10}H_7OH$ C. J. WEST

Anthracene derivatives. II. EDWARD DE BARBY BARTETT AND NORMAN P. GOODWAY *Ber* 63B, 3049-51 (1930), cf. C. A. 24, 5006. — In order to obtain further data on the influence of substituents in the side ring on the reactions of atoms and groups in the *ms* position, attempts were made to reduce 1,5-dimethoxyanthraquinone (I) to the corresponding anthrone, but in spite of numerous trials in both acid and alk. solutions neither the anthrone nor the anthracene could be obtained. The now commercially prepd. and hence readily available 1,5-diphenoxyanthraquinone (II) also did not yield the anthrone (III) by either the usual Sn/HCl or the Al powder-reduced H_2SO_4 method. III was obtained in satisfactory yield, however, with $SnCl_4$ and HCl in $AcOH$. It was unchanged by $MeMgI$ and $PhCH_2MgCl$, the copious evolution of gas with $MeMgI$ pointed to the anthranol rather than the anthrone structure. Attempts to obtain a mono- or di-benzylanthrone by boiling with $PhCH_2Cl$ and $NaOH$ gave only resinous products and unchanged III. III is readily brominated and the 10-Br deriv. (IV) resembles 1,5-dichlorobromanthrone more than bromanthrone itself. It readily yields 10-RO and 10-NHR derivs. This seems to support the theory of peri-coordination, but in view of the contradictory results recently obtained, a theoretical discussion is reserved for a later date. With the object of detg. the influence of a β - PhO group, $C_6H_5(CO)_2O$ was condensed with Ph_2O but the resulting 4'-phenoxybenzophenone 1-carboxylic acid (V) could not be converted into the anthraquinone. It was, however, easily reduced to 4'-phenoxydiphenylmethane 1-carboxylic acid (VI), attempts to convert this into the anthrone have thus far failed. The I, pale yellow, m. 241°, was obtained in about 20 g. yield from 1,5- $NaO_2SC_6H_4(CO)_2C_6H_5Na$ and 1.5 parts $NaOH$ boiled 40 hrs. with continuous stirring in $MeOH$. III, yellow, m. 166°, gives with C_6H_5N - AcO on the H_2O bath 1,5-diphenoxyanthranyl acetate, yellow, m. 165°. IV, yellow, m. 190-3° (gas evolution) 10- EtO deriv., from IV refluxed in alc. with $CaCO_3$, m. 148°, quickly turns light pink on the surface anthranyl acetate, yellow, m. 238°. 10-Piperidino deriv., m. 129°. 10-Anilino deriv., yellow, m. 199° (decompos.). 10-N-Methylamino deriv., yellow, m. 129°. 10-p-Dimethylaminophenyl deriv., faintly yellow, m. 205°. V, m. 162°. VI, from V with Zn dust and aq. NH_4OH on the H_2O bath, m. 123°. C. A. R.

Oxalic acid as a condensing agent. I. Condensation of phthalic anhydride and phenol. S. NAKAMISHI. *J. Pharm. Soc. Japan* 51, 31-4 (1931); German abstr. 10-11 — $C_6H_4(CO)_2O$ (I) (10 g.), when treated with hydroquinone (5 g.) and $(CO_2H)_2$ (20 g.) at 120-40° for 2 hrs., gave griseizarin, $C_{10}H_6O_6$, m. 194°. I (5 g.), α -naphthol (5 g.) and

(CO_2H)₂ (20 g), when treated as above, give α -naphthofluoran, $\text{C}_{18}\text{H}_{10}\text{O}_2$, m 300° (yield 3 g) F I NAKAMURA

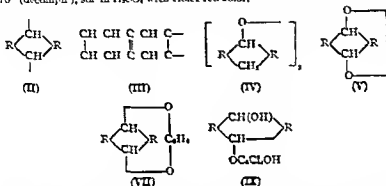
Constitution of the alums. E LEGER *Bull soc chim.* [4], 49, 70-2 (1931).—A polemic. Cf Gibson and Simonsen, *C A* 24, 2752 ODEY E SHEPPARD

Anthraquinhydrone. OSCAR BALLY *Ber* 63B, 2938 (1930), cf Scholl and Böttger, *C A* 25, 695.—In 1905 (Ger pat 191,252), in applying alic KOH fusion to the prepn of isoviolanthrone B observed with anthraquinone itself an intense bottle-green color. Scholl, apprised of this observation, ascribed the color to the formation of K anthraquinhydrone C A R.

Polynuclear aromatic hydrocarbons and their derivatives. VII. A new class of deeply colored radical hydrocarbons and the supposed pentacene of E. Philippi; also a reply to remarks of Roland Scholl and Oskar Böttger. C CLAR AND FR JOHN *Ber* 63B, 2967-77 (1930), cf *C A* 24, 2157.—*lin*-Dibenzo-9,10-dihydroanthracene (I), when dehydrogenated by passing over heated Cu *in vacuo* or by boiling in PhNO_2 with phenanthrenequinone or in xylene with chloranil, gives a deep violet-blue hydrocarbon, $\text{C}_{20}\text{H}_{14}$, to which (see below) is assigned the structure of a 2,3,6,7-dibenzoanthracene 9,10-diyl (II, R = III). The homologs of I yield corresponding violet blue hydrocarbons. These deeply colored very difficultly sol and high melting hydrocarbons differ from the hitherto known condensed aromatic hydrocarbons of the same mol. magnitude not only in their deep color but also in their interesting reactions. The red xylene soln of II is unchanged by dry O but moist O or air changes it to yellow, with green fluorescence, and ppts an insol colorless peroxide (IV), and the mother liquor on evcn yields yellow brown needles, $\text{C}_{20}\text{H}_{12}\text{O}_2$ (V). IV has the compn. $\text{C}_{40}\text{H}_{20}\text{O}_2$ and regenerates II when heated to its m p or in high boiling solvents (PhNO_2), properties agreeing well with the peroxide structure given. IV is then apparently formed by the addn of H_2O_2 , not of O_2 . Unfortunately mol-wt. detns could not be made because of the insolv of the IV. The endocyclic peroxide V is accompanied by varying amts of the monoquinone R(CO)₂R (VI), apparently formed by further oxidation of V. In 1 case where the di mol IV was not formed, the monomeric V was readily isolated in pure form. It does not dissolve in NaOH and on fusion shows the decompn. characteristic of peroxides. Light greatly accelerates the oxidation. CrO_3 and AcOH, as already reported, give VI exclusively. In xylene suspension, II is immediately decolorized by benzoquinone with the formation of a faintly yellow cryst. product to which is assigned the structure VII; on heating it does not break down into its components but melts to a dark mass, with evolution of gas. With chloranil in xylene the reaction is somewhat slower and there are obtained a yellow (VIII) and a brown-red product (IX). VIII, from its compn., is the analog of VII but on heating alone or in high boiling solvents breaks down into its components. IX contains 1 mol more of H_2O and does not split up into its components when heated, it is provisionally assigned the structure shown. Dil. NaOH converts it into a colorless, difficultly sol. compd., while VIII dissolves only in alic. KOH, with brown red color. Crystn. of IX from AcOH changes it into a light red substance (X) of much higher m p, which is no longer altered by dil. NaOH. With phenanthrenequinone, II does not react, it reacts at once with halogens, including I, and readily with boiling PhNH_2 , contg a little $\text{PhNH}_2 \cdot \text{HCl}$. Among the reaction products with halogens are halogen-contg diradicals of great stability. II, which sublimes undecompd. below 300° in CO_2 , is sublimed above this temp., it gives, depending on the temp., small quantities of I, m 270°, and the residue, which retains the cryst. structure of the II, has a glittering, graphite like appearance. It is probably formed by a graphite like linking together of many hydrocarbon mols, while the II thereby set free reduces a part of the II to I. The properties of VI indicate that it has the structure shown. Rontgen ray studies by Debye's method have established the uniplanar position of the C_{10}H_8 rings and leave hardly any doubt as to the correctness of the structure II. The supposed 2,3,6,7-dibenzoanthracene, which Philippi, with Scholl, Houben and de Diesbach, calls *lin* pentacene (*C A* 24, 356), is really the dihydro deriv. (I). Using their own method of prepn (*C A* 24, 1634), C and J obtained a hydrocarbon with the same properties as P.'s but with the compn. $\text{C}_{20}\text{H}_{14}$, not $\text{C}_{20}\text{H}_{12}$, and yielding II on dehydrogenation, and VI exclusively on oxidation. On standing some months, the m. p. of the I rises to 300-10° and the substance (XI) now agrees in all its properties with that first described by Mills and Mills, on oxidation it gives the diquinone, $\text{C}_{20}\text{H}_{12}(\text{CO})_2\text{C}_{10}\text{H}_8(\text{CO})_2\text{C}_{10}\text{H}_8$, although with FeCl_3 it gives only VI; this is readily explained by an intermediate dehydrogenation to II by the FeCl_3 . XI, which is apparently more stable at room temp. than I, is probably formed from I by a change in position of two H atoms and is the 1,4-dihydro

compd., $\text{RC}_6\text{H}_5 \begin{array}{c} \diagup \text{CH}_2 \\ \diagdown \text{CH}_2 \end{array} \text{C}_6\text{H}_5$. The reverse rearrangement has been effected by heating

XI for some time above its m. p. under CO_2 , the resulting product with CrO_3 now gives **VI** along with traces of the diquinone. There is no change in compn., as P. believes, in the rearrangement. Both forms give red melts. A brief reply is made to the criticisms of Scholl and Bottger (*C. A.* 25, 696). **IV** decmps. $535-60^\circ$, imparts a faint green color to concd. H_2SO_4 . **V**, m. $320-30^\circ$ (decompn.), sol. in H_2SO_4 , with green color 2.36; *Dibenzanthracene 9,10-diol hydroquinone ether* (**VII**), decmps. 260° with gas evolution (m. 335° with carbonization if heated very slowly). *Tetrachloro-hydroquinone ether* (**VIII**) m. 263° (gas evolution), sol. in H_2SO_4 , with dirty green color, *mono-ether* (**IX**), m. $207-8^\circ$ (gas evolution), also sol. with dirty green color in H_2SO_4 . **X**, apparently formed from **IX** by loss of 2 mols. HCl and 1 mol. H_2O , darkens 300° , m. 370° (decompn.), sol. in H_2SO_4 , with violet red color.



C. A. R.

Unsaturated compounds. XII. Action of hydroxylamine on fufurhippuric acid and its esters, as also on fufuracrylic esters. THEODOR POSNER AND ILMENGARD STRECKT, *HEM. MONAT.* Ber 63B, 3078-88 (1930); cf *C. A.* 18, 3367.—The addn. at conjugated double bonds has hitherto been studied chiefly on variously substituted cinnamic acids. The present work deals with the influence on the additive power of the replacement of the C_6H_5 nucleus by the furan ring and the relation to the strength of the acid. *Acid addn. products*—The furan ring increases the disson. of the acid and quite materially increases the difficulty of addn., *fufuracrylic acid* (I) adds more difficultly than hippuric acid (II), *fufurhippuric acid* (III) more difficultly than benzalhippuric acid (IV) and III more difficultly than I; the increase in the yield of NH_2OH addn. products of the acids of the furan series agrees well with the decrease in the disson. of the acids, insofar as these constts. have been detd. On short heating the formation of hydroxamino acids, $\text{RCH}(\text{NHOH})\text{CH}_2\text{CO}_2\text{H}$, increases with decrease in disson., and on longer heating the oxamino acids are more and more oxidized to the ketones. The formation of the other oxidation product, the isoxazalone, $\text{N} \begin{array}{c} \diagup \text{CR} \\ \diagdown \text{CH}_2 \end{array} \text{CO}_2\text{O}$, and of

the reduction product, $\text{RCH}(\text{NH}_2)\text{CH}_2\text{CO}_2\text{H}$, is irregular. *Ester addn. products*—The esters add much more easily than the corresponding acids, the Et ester generally somewhat easily than the Me ester, they generally react quite well with NH_2OH in the cold, whereas the acids only form the NH_2OH salts. Unlike the acids, which add only 1 mol. NH_2OH when heated, the esters react in the cold with 2 and 3 mols. NH_2OH and the ability to add 1-3 mols. increases from the ester of the strongest to that of the weakest acid (the Me ester of III adds 1, the Et ester of IV 2, $\text{PhCH}(\text{CHCO}_2\text{Et})_3$ 3 mols.). Most of the cinnamic esters with HO and MeO substituents on the nucleus are exceptions, they in general do not add at all in the cold. The yield of oxime and isoxazalone increases inversely proportionally to the disson. constts. of the corresponding acids, only $\text{o-O}_2\text{NC}_6\text{H}_4\text{CH}(\text{CHCO}_2\text{Et})_3$ is an exception. The amino acids also are generally formed in smaller amounts the weaker the corresponding acids are (except with the esters of III, $\text{o-O}_2\text{NC}_6\text{H}_4\text{CH}(\text{CHCO}_2\text{H})_3$ and *trans-o-MeOC}_6\text{H}_4\text{CH}(\text{CHCO}_2\text{H})_3*). The ester of III forms oxidation and reduction products even in the cold, and in spite of the slow addn. even when boiled, it yields a relatively large amount of amino acid. In the expts. at the boiling temp., esters of III, contrary to the cinnamates, form small

1 g IX with 8 cc concd H_2SO_4 and 2 cc AcOH , m 125-6°. The lack of color is unexpected if the suggested formula is correct (cf Weissgerber and Kruber, *C. A.* 15, 88 who point out that the Na compd of X is not yet explained). Diphenyl 2(?) thionaphthylmethane (XII) $\text{Ph}_2\text{C}(\text{CH}_2\text{C}_6\text{H}_4\text{S})_2$ obtained in 90% yield by splitting out CO_2 from IX as VI was obtained from I, m 101-5°, dissolves in concd H_2SO_4 with intense greenish orange color. Diphenyl-3(?) bromothionaphthylacetic acid (XIII), $\text{Ph}_2\text{C}(\text{C}_6\text{H}_4\text{BrS})\text{CO}_2\text{H}$ obtained in almost quant yield by treating 17 g IX with 16 g Br, m 223-4° (decompn) dissolves in concd H_2SO_4 with violet red color. Diphenyl-3(?) bromothionaphthylcarbinol (XIV) $\text{Ph}_2\text{C}(\text{OH})(\text{C}_6\text{H}_4\text{BrS})$ obtained in 75% yield by the method of prep. XII from IX, m 166-7°, dissolves in concd H_2SO_4 with violet-red color. 4,4'-Dimethoxydiphenyl (?) thionaphthylacetic acid (XV), $(\text{MeOC}_6\text{H}_4)_2\text{C}(\text{C}_6\text{H}_4\text{S})\text{CO}_2\text{H}$ obtained in 90% yield by treating a soln of 2.9 g of anisole acid and 17 g X in 20 cc AcOH with 20 cc of a mixt of equal wts H_2SO_4 and AcOH , letting stand 2 days, m 245-6° (decompn) the Ba salt is almost insol in water, Me ester, in 90% yield prepd as the Me ester of IX, m 166-7°. CO evolution resulted when XV was treated with H_2SO_4 or H_2SO_4 and AcOH but no product was isolated. 9-(2?)Thionaphthylxanthene O $(\text{C}_{12}\text{H}_8\text{S})_2\text{C}(\text{CH}_2\text{C}_6\text{H}_4\text{S})_2$ obtained in 50% yield by refluxing 2 hrs 2 g VIII, 17 g X, 3 g P_2O_5 and 5 cc ether, and letting stand 2 days, m 172-3°, dissolves in H_2SO_4 with yellow color and intense blue green fluorescence. O I S

New methods of formation of coumarone and of diphenylene oxide. N. A. OZOLY AND V. V. TISHCHENKO *Ber* 63B, 2448-51(1930), cf *C. A.* 24, 110—Coumarin (350 g), passed at the rate of 50 g per hr through a 2-cm tinned Fe tube, 60 cm long, at 800° gave a gas consisting of 4.0% CO_2 , 71.6% CO , 19.2% H_2 and 4.8% C_2H_2 , and 297 g of a thick brown product sep'd by distn into 56 g b 80-200° (I), 16 g b 200-48°, 75 g b 288-305° and 140 g residue. The 3rd fraction and a considerable portion of the 2nd fraction consisted of unchanged coumarin and the residue, which was very similar in all its properties to coumarone resin, gave on dry distn 78.6 g oil (II), which was combined with I, and 56 g of a porous coke. The combined I and II, freed of acid products (phenols) with NaOH , was sep'd by careful fractionation into 8 g b 80-172° (chiefly C_{12}H_8 hydrocarbons), 122 g practically pure coumarone, b 172-4°, and 8 g residue (mainly coumarin). Under pressures greater than atm the decompn of coumarin follows another course, heated 4 hrs at 420° in an autoclave the coumarin gave only gases and coke. Coumarin may be considered as an α -benzopyrone, and xanthone (dibenzo- γ pyrone) (100 g) treated like the coumarin at 880°, gave a gas consisting of 94% CO , 1.76% Cl_2 , and 2.1% H_2 , and a black cryst mass giving on distn 15 g of a product b 285-307° which, through the picrate, gave 0.5 g diphenylene oxide, m 85°. Dimethylpyrone was similarly treated, but because of lack of material the best conditions for a smooth degradation into CO and dimethylfuran could not be worked out. The formation of a volatile compd with pronounced furan properties and the evolution of large quantities of gas rich in CO was observed, however.

C. A. R.

The hydrolysis of acetylindoxyllic acid and the acetylindoxyls. GEOFFREY SPENCEA *J. Soc. Chem. Ind.* 50, 63-4T(1931)—*N*-Acetylindoxyllic acid, $\text{AcN C}_6\text{H}_3\text{CO CHCO}_2\text{H}$

(I). A melt of 10 g of *o*-HO.CC₆H₃NHCH₂CO₂H and 45 g of 80% KOH, fused at 260° for 10 min, was ground up, mixed with 50 g of ice and partly neutralized with 35 cc of AcOH at 0° in a CO_2 atm. Ten cc of Ac_2O was added and after 30 min the ppt of O acetylindoxyl was filtered off. Acidification of the filtrate with HCl gave I, m 179° (decompn). I was also obtained by acetylating acid indoxyllic acid solns. Hydrolysis of I was expected to give indoxyllic acid, indoxyl and then indigo but on the contrary a marked tendency to ring cleavage was shown. Hydrolysis with Na_2CO_3 , alk H_2O_2 and NaOH gave 50% of $\text{NH}_2\text{C}_6\text{H}_3\text{CO}_2\text{H}$ and some indigo, whereas treatment with HCl or acid FeCl_3 yielded $\text{NH}_2\text{C}_6\text{H}_3\text{CO}_2\text{H}$ solely. *N*-Acetylindoxyl (II) was prep'd by the method of Vörländer and Pfeiffer (*C. A.* 13, 236), after attempts to decarboxylate indoxyllic acid by boiling had failed. On boiling II in the presence of air with dil Na_2CO_3 , indigo, diacetylindigo and isatin were obtained in 55, 8 and 12% yields, resp. Dil NaOH gave mainly indigo (74%) together with isatin, while acid FeCl_3 gave only isatin. *O*-Acetylindoxyl (III). An indoxyllic acid melt was partly neutralized, boiled for 10 min, cooled, and acetylated with Ac_2O . After standing for 30 min, III was filtered off, washed free from acetylindoxyllic acid and crystalized from H_2O . After drying *in vacuo*, 15% yield of III, m 127.5°, was obtained. Hydrolysis with Na_2CO_3 yielded 95% indigo and 3-4% isatin, treatment with alk H_2O_2 gave mainly isatin, 18% indigo and some $\text{NH}_2\text{C}_6\text{H}_3\text{CO}_2\text{H}$. Acid FeCl_3 gave indirubin (18%) and isatin.

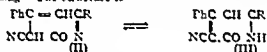
C. R. ADONALD.

3,5-Diketopyrazolidine derivatives. I. TOKUTCHI TSUMAKI *Bull. Chem. Soc. Japan* 6, 1-9 (1931).—1,2-Diphenyl-3,5-diketopyrazolidine (I) $(\text{PhNH})_2$ (18.5 g) in 500 cc of Et_2O was added from a dropping funnel to a soln of 10 g of $\text{CH}_2(\text{COCl})_2$. The white ppt washed with Et_2O in the absence of moist air, was extd with petroleum ether and recrystd from EtOH , yield 6 g, m 175° . The ether filtrate was washed with water and shaken with 1% Na_2CO_3 . On acidification with HCl the alk. soln gave a ppt which when dissolved in CHCl_3 and treated with EtOH yielded crystals of 1,2-diphenyl-3,4-dihydroxy-5-pyrazolidone (II), m $160-2^\circ$. This compd was also obtained by dissolving I in a mixt of CHCl_3 and EtOH , and allowing the soln, acidified with a trace of HCl , to stand for a few days. I and II were obtained by the action of C_6O_3 on $(\text{PhNH})_2$, the poor yields being accounted for by polymerization of the main part of the C_6O_3 during the reaction. I forms an NH_2 salt, m $138-9^\circ$, and an isomeric osone deriv., m $163-4^\circ$. It condenses readily with aldehydes and ketones, giving products

easily decompd by hot alkali. From BzH , O CH CH CH CCHO , PhCH CHCHO , Me_2CO , MeIz and PhIz the following 1,2-diphenyl-3,5-diketopyrazolidine derivs were formed: 4-benzylidene, m $164-5^\circ$; 4-furfurylidene, m $157-8^\circ$; 4-cinnamylidene, m $190-2^\circ$; 4-isopropylidene, m 143° ; 4-(α -phenylethylidene), m $148-9^\circ$, and 4-(diphenylmethylidene), m 29° . C. R. ARDRAVALL

Equilibrium and "intermediate stage." T. ARNDT *Ber* 63B, 2963-6 (1930).—In agreement with v. Auwers *et al.*, who, in their work on the N -alkyl γ -pyridones (C. A. 25, 205), misunderstood the attempted explanation, in footnote 2 to the paper by A. and Kalischek (C. A. 24, 3230) on an electronic theoretical basis, of A's conception of the difference between equil and "intermediate stage" A explains more fully this conception, which is dealt with in detail by his pupil Lorenz (*Dist. Breslau*, 1927) but is referred to only incidentally here and there in A's own papers. C. A. R.

Reactivity of conjugated systems. II. Condensation of acetylenic ketones with cyanoacetamide. CHITTARANJAN BHARAT *J. Indian Chem. Soc.* 7, 651-62 (1930); cf. C. A. 24, 4785.— PhCOC CPh and $\text{CNCH}_2\text{CONH}_2$ condensed according to Michael's method to 3-cyano-4,6-diphenyl-2-pyridone (I), m 320° , in 60% yield. I is identical with the dehydrogenatd product from $\text{PhCOCCH}_2\text{CHPh}$ and $\text{CNCH}_2\text{CONH}_2$. The Na salt of I is sparingly sol in water, on much diln the soln becomes turbid from hydrolysis. With Knoevenagel's reagent 40-50% of an identical compd. was obtained on 1 week's standing. The structure is

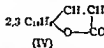
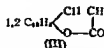
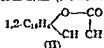


of which III predominates in the presence of alkali. The N -Me deriv (IV), m. 175° , is identical with that from condensation with $\text{CNCH}_2\text{CONHMe}$. Diln of the H_2SO_4 soln of I gives 4,6-diphenyl-2-hydroxypyridine (V). I and II give no color with alc FeCl_3 but V does. V does not yield a Me deriv readily and is probably $\text{PhC CH CR N C(OH) CH}$. V could also be reached from I by an intermediate

compd. $\text{PhC}(\text{CH}_2\text{COR})\text{C}(\text{CN})\text{CONH}_2$. The stability of the condensation products increases with increasing neg. character of R. An extension of Fry's substitution rule is offered in explanation. *p*-Tolylstyrylacetone, bp $200-10^\circ$, m. 72° , and $\text{CNCH}_2\text{CONH}_2$ gave 3-cyano-4-phenyl-6-*p*-tolyl-2-pyridone (VI), m 268° . VI was also obtained from $\text{p-MeC}_6\text{H}_4\text{COCH CHPh}$ and from α -methoxy- α -*p*-tolylstyrene; N -Me deriv., m $136-8^\circ$. 4-Phenyl-6-*p*-tolyl-2-hydroxypyridine, m 228° . Et 4-phenyl-6-*p*-tolyl-2-pyridone-3-carboxylate, m 85° . 4-Phenyl-6-*p*-tolyl-2-pyridone, m 133° . *p*-NO₂-C₆H₄COCHCHPh, m $161-2^\circ$. 3-Cyano-4-phenyl-6-*p*-nitrophenyl-2-pyridone, m $342-3^\circ$; N -Me deriv., m $322-4^\circ$. 4-Phenyl-6-*p*-nitrophenyl-2-hydroxypyridine, m $275-6^\circ$. Et 4-phenyl-6-*p*-nitrophenyl-2-pyridone-3-carboxylate, m 145° . 4-Phenyl-6-*p*-nitrophenyl-2-pyridone, m 235° . 3-Cyano-4-phenyl-6-methyl-2-pyridone, m $275-6^\circ$; N -Me deriv., m 146° . 4-Phenyl-6-methyl-2-hydroxypyridine, m $206-8^\circ$. Et 4-phenyl-6-methyl-2-pyridone-3-carboxylate, m 91° . 4-Phenyl-6-methyl-2-pyridone, m $155-6^\circ$. 3-Cyano-4-phenyl-6-ethyl-2-pyridone, m $267-8^\circ$; N -Me deriv., m 188° . 4-Phenyl-6-ethyl-2-hydroxypyridine, m 165° . FOSTER DES SNEILL

The third α -hydroxynaphthaldehyde and the isomeric β -naphthoquinone. THEODOR BOEHM AND ELMAR PROFF, *Arch. Pharm.* 269, 25-37 (1931).—Of the 3 possible α -hydroxynaphthaldehydes only 2 were known, to which the structures 1,2- and 2,1-C₁₀H₇(OH)CHO were ascribed. The prepn. of the 2,3-isomer (I) was un-

successfully attempted some years ago by Zetzsche and co-workers by the catalytic reduction of $C_{11}H_7(OH)COCl$. These expts. were taken up anew by B and P. who, following the procedure of Rosenmund and starting with $2,3\text{-}C_{11}H_7(OH)CO_2H$ succeeded in prep. I, by the following transformations $C_{11}H_7(OH)CO_2H \rightarrow C_{11}H_7(OAc)CO_2H \rightarrow C_{11}H_7(OAc)COCl \rightarrow C_{11}H_7(OAc)CHO \rightarrow I$. This synthesis clarifies the questions hitherto existing in the constitutional make up of the isomeric β -naphthocoumarins (III and IV)



Owing to its derivation from α -naphthol no question arises respecting the structure of α -naphthocoumarin (II). Of the isomeric β -naphthocoumarins, the one prepd by Kauffmann from β -naphtholaldehyde according to Perkin, m. 118°, is represented by (III). A 2nd β -naphthocoumarin was described by v. Pechmann and Welsh (cf. *Ber* 17, 1646), it m. 141° and was designated as *iso-β*-naphthocoumarin. Its constitution was left undetd., from K's work only structure IV would remain for it. However, when β -naphtholaldehyde was subjected to the Perkin synthesis, a product m. 163°, resulted, entirely unlike Pechmann's *iso-β*-naphthocoumarin, and undoubtedly having the structure IV. Repetition of the v. Pechmann expts on condensation of β -naphthol with malic acid in the presence of H_2SO_4 yielded a product m. 118° and identical with K's β -naphthocoumarin. If the same reaction is applied to α -naphthol—which according to Pechmann does not react at all—a product, m. 141°, results (already observed by Bartsch, *Ber* 36, 1966), it is II, which had been previously prepd by another procedure by Bezdruk and Friedlander (cf. *Monatsh Chem* 30, 280). Accordingly, the assumption is safe that v. Pechmann and Welsh had in hand II and that an *iso-β*-naphthocoumarin, m. 141°, has no existence in fact and should be deleted from the literature. During the present study, the following compds. were prepd and characterized: $2,3\text{-}C_{11}H_7(OH)COCl$, m. 100° (Graff, *Ann*, 367, 253, reports m. p. 89°), which gave on reduction *2-acetoxynaphthol aldehyde*, $C_{11}H_{10}O_2$, bright yellow, m. 100–1°, likewise the related by product *acetoxynaphthol anhydride*, $C_{11}H_{10}O_3$, m. 156°. I, $C_{11}H_{10}O_2$, yellow plates from dil. $AcOH$, m. 69–100°, *phenylhydrazones* $C_{11}H_{10}ON_2$, yellow, m. 216–8°, *p*-nitrophenylhydrazones, $C_{11}H_{10}ON_2$, brownish yellow, m. 271–3°, *semicarbazones*, $C_{11}H_{10}ON_2$, m. 270°, *oxime*, $C_{11}H_{10}ON$, m. 207°; *anil*, $C_{11}H_{10}ON$, golden, m. 153–9° (*p*-nitrophenylhydrazones, $C_{11}H_{10}ON_2$, golden, m. 213–4°, *semicarbazones*, $C_{11}H_{10}ON_2$, m. 211–2°, *oxime*, $C_{11}H_{10}ON$, m. 202–3°), β -Hydroxy- β -naphthylidenemalonate, $C_{11}H_{10}O_4$, yellow, m. 207–8°, β -Hydroxy- β -naphthylidenemalonate, $C_{11}H_{10}O_4$, yellow, m. 183–9°, β -Hydroxy- β -naphthylidenemalonate, $C_{11}H_{10}O_4$, bright yellow, m. 162–3°, β -Hydroxy- β -naphthylidenemalonate ester, $C_{11}H_{10}O_4$, yellow, m. 112–3° (*Ac* deriv, $C_{11}H_{10}O_4$, m. 142–3°). Condensation of I with AcO and $AcONa$ yielded IV, yellowish white, m. 163–4°. *Alc.* $PhNH_2$ or NH_3 effects condensation of I with malonic acid to yield *iso-β*-naphthocoumarincarboxylic acid, $C_{11}H_{10}O_5$, yellow, m. 253–9° (*Et* ester, $C_{11}H_{10}O_5$, yellow, m. 157–8°). With $AcCH_2CO_2Et$ and piperidine I yields *3-acetyliso-β*-naphthocoumarin, $C_{11}H_{10}O_4$, yellow, m. 209–10° (*p*-nitrophenylhydrazones, $C_{11}H_{10}ON_2$, yellow, m. 266–7°, *Et* deriv, $C_{11}H_{10}O_4$, yellow, m. 223°). *Et* (3-cyano-3,4-dihydro-*iso-β*-naphthocoumarin)cyanoacetate, from the condensation of I with $NCCH_2CO_2Et$ in the presence of piperidine, m. 153–9°.

W. O. E.

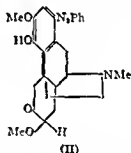
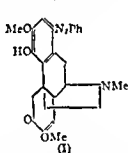
α -Naphthoic acid. HENRY GILMAN, NINA B. ST. JOHN AND F. SCHULZE. *Org. Syntheses* XI, 80–3 (1931).—Complete details are given for the prep. of α - $C_{11}H_7CO_2H$ from α - $C_{11}H_7Br$ through the Grignard reagent, yield, 68–70%. C. J. West

10-Chloro-5,10-dihydrophenarsazine and its derivatives. XIV. Chloro derivatives. LESLIE A. ELSON AND CHARLES S. GIBSON. *J. Chem. Soc.* 1931, 294–305, cf. C. A. 24, 5301.—Condensation of Cl derivs. of Ph_2NH with $AsCl_3$, producing derivs. of 5,10-dihydrophenarsazine, takes place except when the Cl is in the *o*-position. The production of isomers from *p*- $YC_6H_4NHC_6H_4X$ and from *o*- $YC_6H_4NHC_6H_4X$ is discussed and their orientation considered. 2,2'- $Cl_2C_6H_3NHC_6H_4CO_2H$, heated 5 hrs. at 300–20°, gives *o*-chlorodiphenylamine, b_p 174–5°, m. 15–6°, 2,5- Cl_2 deriv, b_p 195°, does not solidify at –16° (85% yield), 3,4- Cl_2 deriv, b_p 218–20°, m. 69°, in 75% yield from 3,4-dichlorodiphenylamine-2'-carboxylic acid, pale yellow, m. 178°, 3,5- Cl_2 deriv, b_p 200°, does not solidify at –16°, in 74% yield from the corresponding 2'-carboxylic acid, m. 245°. 3,4,5- Cl_3 deriv, b_p 248°, m. 85°, from the 2'-carboxylic acid, m. 238°. 5-Chloro-2-methyldiphenylamine, b_p 205–7°, from the 2'-carboxylic acid,

pale yellow, m 200° *N*-3-Chlorophenylbenzazino-4-chlorophenyl ether, m. 77°, from benz-*m*-chloroanilide aminochloride and $p\text{-ClC}_6\text{H}_4\text{ONa}$, heating 2 hrs at 290° and hydrolysis of the Bz deriv. gives 3,4'-dichlorodiphenylamine, b_{11} 235°, m 63-4°, 3,3'-Cl₂ deriv, b_{11} 225-30° *N*-o-Tolylbenzazino-2-chlorophenyl ether, m 64°, heating 2 hrs at 290° gives benzoyl-2'-chloro-2-methyldiphenylamine, m 132°, which could not be hydrolyzed by the usual methods, the corresponding 3'-Cl deriv m. 106°; hydrolysis gives 3'-chloro-2-methyldiphenylamine, b_{11} 203° *N*-4-Chlorophenylbenzazino-o-tolyl ether, m 65°, isomerization and hydrolysis give 4'-chloro-2-methyldiphenylamine, b_{11} 200-5°, darkens in the air *N*-o-Tolylbenzazino-2,4-dichlorophenyl ether, m 53° this gives 2',4'-dichloro-2-methyldiphenylamine, b_{11} 205°, m 48° Condensation of 3,4 Cl₂C₆H₃NHPh and AsCl₃ in $o\text{-C}_6\text{H}_4\text{Cl}_2$ gives 20% of a mixt of 1,2,10- and 2,3,10 trichloro-5,10-dihydrophenarsazine, m 265-8° and 220°, the same compds were obtained by reducing I with SO₂ in HCl, the products m 273-4° and 230-5° 3,4 Cl₂C₆H₃NH₂, $o\text{-BrC}_6\text{H}_4\text{AsO}_2\text{H}_2$, K₂CO₃, ArOH and Cu boiled 8 hrs, give 3,4-dichlorodiphenylamine-6'-arsonic acid (I), m 150° 3,5 Cl₂C₆H₃NHPh and AsCl₃ give the 1,3,10 Cl₂ deriv, yellow, m 251-2°, boiling with 20 vols H₂O and AcOH gives 1,3-dichlorophenarsazinic acid, does not m 280° 3,4-5-Cl₂C₆H₃NHPh and AsCl₃ give 1,2,3,10 tetrachloro-5,10-dihydrophenarsazine, yellow, m 260°, oxidation gives 1,2,3 trichlorophenarsazinic acid, does not m 290° (3 ClC₆H₄)₂NH gives 1,9,10-, 1,7,10 or 3,7,10 trichloro-5,10-dihydrophenarsazine, yellow, m 298° 3,4' ClC₆H₄NHCl and AsCl₃ give a mixt. of the 1,8,10- and the 3,8,10 tri Cl derivs, m 292° and 240 2°, resp, the yield of the mixt. is about 20% 4,2 ClMeC₆H₃NHPh and AsCl₃ give 70% of 2,10-dichloro-4-methyl-5,10-dihydrophenarsazine, bright yellow, m 245° (decompn), 3 ClC₆H₄NHCl₂ H₂Me-2 gives 53% of the 1(or 3),10-dichloro-6-methyl deriv, bright yellow, m 262° (decompn), 4 ClC₆H₄NHCl₂H₂Me-2 gives 25% of the 2,10-dichloro-6-methyl deriv, deep yellow, m 199°

C J WEST

Sinomenine and dislaomenine. XX. Benzeneazosinomenine. KAKUJI GOTO AND HIDEO SUISHIDO *Bull Chem Soc Japan* 5, 311-4(1930), cf. C A 25, 299 —By means of a diazo coupling reaction, a radical is introduced in the sinomenine mol, in the β -position with regard to the phenol group Benzeneazosinomenine (I), dark red tetragonal or hexagonal plates, m 253° (decompn) Benzeneazodihydrosinomenine (II), red prisms m 231° (decompn) Benzeneazothedainone, m 152° Ammodihydro sinomenine HCl salt, m. > 300°.

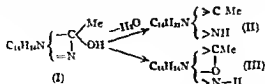


XXI. The reaction between sinomenine and formaldehyde. KAKUJI GOTO, HIDEO SUISHIDO AND REIKICHI INADA *Ibid* 315-20 —The interaction of sinomenine and 40% HCHO gives 5-hydroxymethylsinomenine (I) and 1,5-dihydroxymethylsinomenine (II) 2 m 280° (decompn), $[\alpha]_D^{20}$ -40.71°, oxime, m 240-5°, methiodide, m 223°, gives a green FeCl₃ reaction, a faint ferricyanide reaction, an intensely red diazo reaction and a green brown HCHO-H₂SO₄ reaction. II m. 242° (from alc) or 252° (from CHCl₃), $[\alpha]_D^{20}$ -74.39°, oxime m 200-15°; methiodide m. 210°, decompn 280°, gives a brown FeCl₃ reaction, a red diazo reaction, no ferricyanide reaction and a green HCHO H₂SO₄ reaction. The reduction of V gives dihydro-5-hydroxymethylsinomenine, m 244°, $[\alpha]_D^{20}$ 73.03°; oxime, m 215-25° (decompn), gives a green FeCl₃ reaction, an intensely red diazo reaction, a faint ferricyanide reaction and a blue-violet HCHO-H₂SO₄ reaction

ALBERT L. HENNE

Constitution of matrine. XIII. Action of Grignard reagent on matrine (I). H KONDO, E. OCHIAI and K. IWAYA. *J. Pharm Soc Japan* 51, 1-8(1931), German abstr. 1-3; cf. C. A. 23, 2447 —Matrine (18.5 g), when treated with MeMgI, followed by decompn with NH₄Cl, gave a compd (I), m 221-30°, with mol formula of C₁₁H₁₇N₂O III H₂O or C₁₁H₁₅N₂O III H₂O. On hydrolysis I gave bases (II), b_1 165° and

(III), b_p 236–40° II, m 65°, d_4^{20} 1.0342, n_D^{20} 1.531866, picrate decomps. 268°; atyphinate decomps. 306° II cannot be reduced with a catalyst, but gives pos diazo and Liebermann reactions II can be neither benzoylated nor acetylated III, $C_{11}H_{11}N_2O$ or $C_{11}H_{11}N_2O$, gives pos diazo and Liebermann reactions, but the reaction for the CO group is neg The reaction is shown in the following equation.



F. I. NAKAMURA

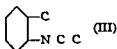
Mechanism of the oxidation of laudanone. VINAYAK K. BHARGAVAT, DOUGLAS K. MOORE and FRANK L. PYMAN *J Chem Soc* 1931, 443 —The compd., $C_{11}H_{15}(OMe)_2$, m 231°, obtained by Pyman in the oxidation of laudanone by MnO_2 in H_2SO_4 (*C A* 3, 2666) is 2,3,6,7-tetramethoxy-9,10-dihydroanthracene (Robinson, *C A* 9, 1602), it seps from Me_2CO with 0.5 H_2O and from CaH_2 with 0.5 CaH_2 . This identification counteracts the view of Gadamer (*C A* 10, 1033) of the mechanism of this oxidation.

C. J. WEST

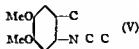
Oxonitine and remarks on a paper of S. Akabori and K. Saito. ERNST SPÄTH AND FRIEDRICH GALINOVSKY *Ber* 63B, 2997–7(1930) —Earlier workers have assigned to oxonitine a formula with 23–25 C atoms, with 3 MeO groups, and it therefore seemed that if the 3 MeO, the Ac and the H_2 groups were split off there would remain a compd. with only 11–13 C atoms whose structure, it was hoped, might be established and prove of great importance for the elucidation of the chemistry of the acetonitines. It was soon found, however, that the formulas of the earlier investigators are incorrect, as oxonitine contains only 2.21% N instead of the 3.04–2.76% required by these formulas, and 4 instead of 3 MeO groups. Especially convincing evidence of the incorrectness of these formulas was afforded by a careful detn. of the H_2 content by sapon. the oxonitine with alc. KOH, extg. the BrOH with Et_2O , subliming it at 130–50° under 10 mm and weighing the sublimate (a control detn. on acetonitine gave 16.09–16.16% H_2 , calcd. 16.27). In this way oxonitine was found to contain 15.08–16.28% H_2 instead of the 4.1–6.1% calcd. for the older formulas. Numerous analyses gave an av. C and H content of 60.48 and 6.90%. The formula for oxonitine is therefore $C_{11}H_{14}O_3N$ (the true H content may differ from this by 2 atoms) and the hope of obtaining a relatively simple product by its hydrolysis proved to be illusory. Akabori and Saito (*C A* 25, 700), in referring to a paper by Späth and Lederer, convey the erroneous impression that S and L considered their β (6-methoxy- β -indolyl)ethylamine, m 112–30°, to be the pure base, it was obtained by a ring closure proceeding in 2 directions, which would naturally give impure end products, and S and L made no attempt to sep. the resulting mixt. of free bases but isolated them as their N-Ac derivs. C. A. R.

Constitution of the strychnos, yohimbe and quebracho alkaloids. ERNST SPÄTH AND HERMANN BRETSCHNEIDER *Ber* 63B, 2997–3006(1930) —Oxidation of strychnine with alk. $KMnO_4$ on the H_2O bath until the consumption of the $KMnO_4$ slowed up gave a mixt. of acids sepd. by their different acidities, one fraction gave an Et_2O -sol. acid (I) whose purification was rendered difficult by the fact that H_2O , the best solvent for it, gradually decomps. it at high temps. It was accordingly treated with CH_3N_3 , which gave a pure di-Me deriv. $C_{11}H_{11}O_4N$ (II) which regenerated I with alc. KOH. Thermal decompn. of the K salt of I gave o - $H_3NC_6H_4CO_2H$ and I was identified as o - $H_3CCONHC_6H_4CO_2H$ and II as its di-Me ester by synthesis. A sample of the acid, m . 195°, obtained by Warnat by oxidation of yohimboic acid and to which he assigned the compn. $C_{11}H_{11}O_4N$, also proved to be identical with L. Intending to oxidize yohimboic acid themselves, S and B. used for its prepn. a sample of com. yohimbin hydrochloricum and found that the decompn. products and rotation of this base and the acid obtained from it by sapon. agreed almost completely with Hahn's data on quebrachine, and oxidation of the acid obtained from the quebrachine HCl yielded L. These results show that strychnine and the yohimbe and quebracho alkaloids contain the grouping III; whether the N is present in a quinoline or an indole ring cannot as yet be stated. Brucine uses up $KMnO_4$ more rapidly than strychnine and the yield of resulting acid is smaller, this was identified by synthesis as 4,5-dimethoxy-N-oxalylanthranic acid (IV), showing that brucine contains the grouping V, as assumed by Leuchs and by

Perkin and Robinson I, prepd. by Kretschy's method, decomps 229–30° (the m. ps 180–210°, recorded in the literature, are probably due to the presence of impurities acting catalytically) II, m. 152.5°. The *di-Me ester*, m 205–6°, of IV was synthesized by fusing 2,4,5- $\text{I}_3\text{N}(\text{MeO})_2\text{C}_6\text{H}_2\text{CO}_2\text{Me}$ with anhyd. $(\text{CO}_2\text{H})_2$ and methylating the resulting *mono-Me ester*, decomps 230° to a turbid liquid which clears up at 232°, with CH_3N_3 .



(III)



(V)

C A R

The melting point of pure yohimbine (remarks on the communication of G. Hahn and W. Schuch). K. WARNAT *Ber* 63B, 2959–61(1930), cf H and S. C. A. 24, 5039—H and S claim that *pure* yohimbine, obtained by fractional crystn. of a compound containing as chief impurity the so-called α -yohimbine, m 215–6° instead of 230–5°, as generally accepted W, in earlier work on the sepn of these 2 alkaloids, had found that this is by no means easily effected by fractional crystn, for his isoyohimbine is none other than H and S's α -yohimbine, and this isoyohimbine easily depresses the m p of yohimbine to 212–8°. A few new expts with pure yohimbine have confirmed the old m p, 234°, of the base, which showed $[\alpha]_D^{20} 93.8^\circ$ ($\text{C}_6\text{H}_7\text{N}$, c 0.921), HCl salt, m. 302°, $[\alpha]_D^{25} 103.3^\circ$. Yohimboic acid, m 259°, anhydride, ro 298°. The rotation of 70–95° given by H and S for yohimbine-HCl and the decompn temp 284° for the yohimboic acid also point to contamination with isoyohimbine.

C A R

Identity of α -yohimbine with isoyohimbine. G. HAHN and W. SCHUCH. *Ber* 63B, 2961–2(1930), cf C A 24, 5039—Whether or not Warnat's isoyohimbine (see preceding abstr.), which he did not sharply characterize, is identical with α -yohimbine, there is no doubt that the alkaloid decomp 239–40°, which H and Brandenburg called isoyohimbine (C. A. 21, 1815) is different from α -yohimbine.

C A R

Properties of nicotine and its derivatives. III. Chloronicotine and methylnicotine. THOMAS M. LOWRY and HAMID KILAN GORE. *J. Chem. Soc* 1931, 319–23, cf. C. A. 23, 5186—Chloronicotine (I) in EtOH shows an absorption band with max $\log \epsilon = 3.55$ at 2700 A. U., as compared with $\log \epsilon = 3.37$ at 2600 for nicotine. Methylnicotine (II) gives a max. $\log \epsilon = 3.44$ at 3100 A. U. in cyclohexane and 3.66 at 2950 in H_2O . The rotatory dispersion of I–II can be expressed by 1 term of Drude's equation with characteristic wave length $\lambda_c = 2550$ A. U. II gives $\lambda_c = 2668$ A. U. in cyclohexane and 2654 in H_2O . Both compds, therefore, behave like the parent base and not like its salts, which are *d* rotatory and of much lower dispersion.

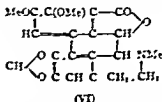
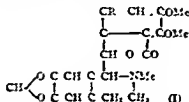
C. J. WEST

Santonine acid amide. KARL JOSEPHSON. *Swensk Farm. Tids.* 35, 69–75(1931); cf. C. A. 25, 1534.—Five g. santonin in 50 cc. abs alc. was satd with dry NH_3 and shaken for over 1 day and then set aside for 2 days in a dark place. The NH_3 was removed by evapn at reduced pressure. Yellowish red crystals formed but the color was completely removed by acetone. The yield is quant. The product m 177°, $[\alpha]_D^{25} -18.1^\circ$, it is more sol in water than santonin and is not very sol. in org solvents; of these alc. is best. On heating it reverts to santonin.

A R. ROSE

Synthesis of hydrastine. I. EDWARD HOPE, FRANK L. PYMAN, FREDERICK G. P. REMFRY and ROBERT ROBINSON. *J. Chem. Soc.* 1931, 236–47, cf. C. A. 7, 2927.—Condensation of hydrastinine and nitroreconin in boiling EtOH gives 80–90% of a mixt. of anhydrohydrastinine and nitroreconins (I, R = NH_2), it is orange-yellow, two samples m. 167–8° and 173°, HCl salt, yellow, m 137–8°, sepn by means of the HCl salt was not successful. Reduction with Sn or SnCl_2 in HCl and AcOH gives a mixt. of aminohydrastines a(II) and b(III), which were sepd. by the difference in soly of the HCl salts and also of the bases. II, m. 216–7° (cor), sparingly sol in EtOH and Et_2O , in which it exhibits an intense blue fluorescence, HCl salt, m. 267–8°, tartrate, m. 201–3° (cor); Ac deriv, m. 208–11°. III, m. 196–7° (cor), cryst. salts could not be obtained. Reduction of the diazo compd from II gives hydrazinohydrastine-a(IV), m. 175° (cor); acid oxalate, m 225° (decompn), isopropylidene deriv, m. 190°. The *b*-isomer (V), crystals with 1 mol CHCl_3 , m. 183–4°, isopropylidene deriv, m 217–8°. Oxidation of IV with $\text{Cu}(\text{OAc})_2$ gives hydrastine a, m 137–8° (cor), HCl salt, decomp. above 165°; acid oxalate, sparingly sol in cold H_2O , picrate, bright yellow, m 219° (decompn.) (the picrate of *b*-hydrastine, bright yellow, m 184°), *H* *d*-tartrate, crystals with 3.5 mols. H_2O , m 108–10° (cor, decompn), *d*-camphorsulfonate, crystals with 0.5 mol. H_2O , m 145° (cor, decompn). V gives only a small quantity of hydrastine-*b*,

m. 154-5° or tartrate, m. 65-70°, the main product of the reaction is *dehydrohydrastine* (VI) m. 152° the H_2SO_4 soln. changes from colorless to yellow, orange, reddish brown, green, to brown, dirty blue and violet on heating. A comparison is made of the behavior of hydrastine, hydrastine-a and -b and VI on oxidation in H_2SO_4 with NaCl , IV and CuCl_2 in about 0.5% HCl give *dehydrohydrastine-a*, m. 152°, reduced to 1.5% hydrastine-a (see p. 215), the I derivative (VII) m. 172° and gives a better yield of hydrastine-a on reduction. Combination of hydrastine and valerianolone gives as the main product VII but some of the base, m. 193-5°, is also formed, the HCl soln. of the 2 isomers m. 222-3° (see, description). Bromination gives a monobromide m. 170-1° (see), HBr soln. m. 2-7-8° (description). Normococaine Me ester with AcCl and Ac_2O gives 50% of *4-acetoxy-3-methoxycinnolide*, m. 127-8° or *4-bromococaine ester* m. 125-6° (see), *methoxycinnolide ester*, m. 122-3° or *paracetamol* m. 232-3°. These acyl deriva. did not condense with hydrastine under any conditions tried. The Me ester itself gives a hydrastinium salt¹, m. 234° or, gives on color with FeCl_3 but on warming with dil. acid or alkali, the base ester of the Me ester could be obtained.



C. J. WEST

Salutaridin of Salix cortex L.; its identity with piceoidin. J. PARATE. *Bull. ex. cas. bot.* 12, 449-73 (1939), of C. A. 24, 4517, 475a. The 2 compounds are identical (3-gluconide of polyhydroxyphenol).

C. G. KING

Synopsis of *Parasitopsis*, Kuhn. V. *Parasitopsis*, S. Aoyama. J. Pharm. Soc. Japan 59, 1152-74 (1939). German abstr. 153-7; C. A. 25, 1533. In the previous

paper A proposed the structure $-\text{CHMe} \cdot \text{CH}(\text{OH}) \cdot \text{CH} \cdot \text{CHMe} \cdot \text{C} \equiv \text{C} \cdot \text{C} \cdot \text{C} \cdot \text{CH}(\text{CO}_2\text{H}) \cdot \text{OH}$ for *parasitopsis* (I) and showed that OH can be acetylated while OH^2 cannot, and that the positions of the unsat. bonds are as shown above. I (50 g.) when treated with an equal wt. of Se at 270-65° for 70 hrs. gave a compd. (II) m. 279-320° which reduced the m. p. of I to 277°. II was fractionated into (A) by 70-100° (0.25 g.), (B) by 130-51° (0.28 g.), (C) by 150-70° (4.45 g.), (E) was distd. further over Na which gave (B) by 125° (1.22 g.), (B) by 125-45° (2.45 g.) and (B) by 145-51° (1.25 g.). The p-oxide of (B) m. 125-7° and corresponds to *isophthalic p-oxide*, $\text{C}_{10}\text{H}_8\text{O}_2$. On CaH_2 over Na (C) gave (C) by 145-49° (0.27 g.) and (C) by 149-54° (2.63 g.). P-oxide of (C) m. 125-7° and corresponds to *terephthalic p-oxide*, $\text{C}_{10}\text{H}_8\text{O}_2$. 1 (C g.) when treated with 5% HCOOH (d. 1.215) for 2 hrs. gave a compd., $\text{C}_{20}\text{H}_{28}\text{O}_6$ (III), m. 243-3.5°, $[\alpha]_D^{25} 65.54^\circ$. III is called *repsin*. The ester of III m. 151-4°, $[\alpha]_D^{25} 22.22^\circ$, *picrylacetate* m. 279°. III cannot be acetylated. Fusion of III with alkali, followed by careful acidification, gave *isorepsin*, m. 233-4°, probably a lactone. The formation of a lactone from I may be due to the formation of an α hydroxy- γ lactone, followed by dehydration.

F. I. NAKAMURA

Acid- and alkali-binding power of peptides. E. STAMM AND H. SCOTTL. *Er. 63B*, 2077-82 (1935). The object of this work was to det., by a study of the titration curves of glycine and its mono-, di-, tri-, tetra- and pentapeptyl deriva. (which contain only 1 NH_2 and 1 COOH group but an increasing no. of peptide unions), whether the peptide unions play a part in the acid and alkali-binding power of peptides. The titration curve of glycine zwitterion (dipotassium) was also detd. The curves show that glycine behaves like a weak monobasic base and a weak monobasic acid. Addition of HCl at first greatly decreases the pH value of its soln. and then slowly decreases the pH value of the resulting buffer. After the addition of 1 equiv. HCl the curve coincides with that of HCl itself. Addition of NaOH increases the pH value in an analogous way. The dipeptide behaves exactly like glycine with HCl but with NaOH the curve does not show an inflection as soon as 1 equiv. NaOH has been added and further addition of NaOH gives lower pH values than would be obtained if the curve coincided with that for pure NaOH . The same behavior toward HCl and NaOH is shown by the higher peptides except that the greater the no. of peptide unions present the further away from the

point corresponding to 1 equiv NaOH is the point of inflection of the curve shifted, and the greater is the difference between the curve for the peptide and that for pure NaOH (the hexapeptide curve is not entirely comparable with those of the other peptides because a suspension had to be used). These results show that the peptide unions take part in the binding of alkali but not in that of acid. The curve for glycine anhydride indicates that there is no acid binding and only a very slight binding of alkali. The following dissociation constants ($K_a \times 10^3$ and $K_b \times 10^{13}$) and isoelectric points (p_{m}), resp. are calcd. for the peptides from the curves: di, 0.63, 1.25, 5.66, tri, 1.0, 1.26, 5.50 tetra, 1.78, 1.12, 5.40, penta, 2.0, 1.12, 5.38 hexa, 2.5, 1.12, 5.32. For glycine, the values are $K_a / 10^3$ 1.99, $K_b / 10^{13}$ 2.6, p_{m} 6.06. C. A. R.

The bile acids. XXXIII. Bromination of dehydrocholic acid and of dehydrodesoxycholic acid. HEINRICH WILLAND and TAKAHASHI NOGUCHI. *Z. physiol. Chem.* 194, 218-59 (1931), cf. *C. A.* 25, 1535.—The 2 simplest keto acids of the bile acid series, viz., dehydrocholic acid (I) and dehydrodesoxycholic acid (II), were brominated and the Br removed by hydrolysis, in order to obtain material for a study of oxidative breakdown. Contrary to expectation, the substitution of Br in all positions adjacent to CO is not easily accomplished. Ring I appears to be the most reactive, since bromination is much more difficult in the case of desoxybiliary acid where this ring has been opened by oxidation. Bromodehydrocholic acid, m. 182-3° (decompn.), was obtained from I and the calcd. quantity of 4 *M* Br in AcOH at 15°. Ester m. 192° (decompn.). Treatment of the acid with cold 0.1 *N* KOH converted it to hydroxydehydrocholic acid, m. 187° (decompn.). Further bromination of I in the cold gave α -dibromodehydrocholic acid which contains solvent of crystn. not removable at 100° in *vacuo*. With 1 AcOH the m. p. is 200°, with 0.5 Et₂O it is 174° and with 1 EtOH the decompn. point is 205°. Treatment of the di Br deriv. with 0.5 *N* KOH gave dihydroxydehydrocholic acid, 2 fractions m. 195° and 212-3°, resp., both contg. 1 Et₂O. A similar treatment with 0.315 *N* Ba(OH)₂, followed by acidification, and Et₂O extn. of the ppt. gave the unsatd. hydroxytriketocholic acid, m. 257° (decompn.), crystg. with 0.5 AcOH. When I was brominated at higher temps with excess Br₂, an isomeric β -dibromodehydrocholic acid, m. 205-6°, was obtained. It holds 1 AcOH tenaciously. This isomer probably results from a rearrangement due to a reversal of the reaction after the 1st Br is substituted. Treatment of I with 6 Br₂ at 30° gave a tetrabromodehydrocholic acid, m. 213°. The mother liquor contained a tribromodehydrocholic acid, decompn. 202-3°, which could not be freed from admixt. with the tetra deriv. Finally, a pentabromodehydrocholic acid, m. 192°, was obtained from I and 8 Br₂. An improved method is described for the prepn. of II, which consists essentially in dissolving I in a large vol. of AcOH, cooling the soln. to 10° and turning with CrO₃ at this temp., then pptg. the product with water, dissolving in Na₂CO₃, reprecip. with HCl and extg. with Et₂O. Bromination of II yielded tribromodesoxycholic acid, m. 212-3° (decompn.), contg. 0.5 AcOH. Recrystd. from MeAc, the acid decompn. 184°. An isomeric tribromodiketocholic acid, m. 213-4°, was obtained from the mother liquor.

A. W. DOX

The biase of chitin. MAX BERGMANN, LEONIDAS ZERVAS and EPHIM SILBERKWEIT. *Naturwissenschaften* 19, 20 (1931).—By acetolysis of lobster chitin a disaccharide of glucosamine was obtained as a cryst. octa Ac deriv., m. 289°, $[\alpha]_D^{20}$ 50° in AcOH, 1 no (cf. B. and Machemer, *C. A.* 24, 2877) 31.3 (calcd. 29.6). The new disaccharide, chitobiose, may be as useful for chitin study as cellulose for cellulose research.

B. J. C. VAN DER HOEVEN

Cholesterol. XIII. Action of phosphorus pentasulfide. E. MONTIGNIE. *Bull. soc. chim.* [4], 49, 73-5 (1931), cf. *C. A.* 25, 1256.—Cholesterol (I) in CS₂ was refluxed with a like wt. of P₂S₅ and yielded thiocholesterol, C₂₇H₄₈SH (II), m. 191°, mp -39°, (di-Br deriv. m. 152-3°, nitro deriv. m. 125°). II does not combine with HgCl₂ and does not give the Liebermann reaction. I reacts with white P and O₂ to yield a non-cryst. phosphate according to the equation: $\text{—CH}_2\text{—P(=O)(OH)—CH}_2\text{—} + \text{P}_4 + 2\text{O}_2 \rightarrow$

$\text{—CH}_2\text{—P(=O)(OH)—P(=O)(OH)—CH}_2\text{—}$. Upon hydrolysis of this compd. (OH)₂P(=O)OCH₂CH₂—

PH(=O)OH was obtained. The Ag, Pb, Ba, Sn⁺⁺, Zn, Hg, U, Cu and Hg salts were prepd. The acid yields an acetylated product, m. 250°, decompn. about 200°. The OH group is left intact.

I. M. LEVINE

Ursolic acid. CHARLES E. SANDO. *J. Biol. Chem.* 90, 477-95 (1931).—Ursone (I) was isolated from the leaves of *Arctostaphylos uva-ursi* by extg. with Et₂O, dissolving in EtOH contg. 1% NaOH and repeatedly crystg. from aq. EtOH contg. a little NaOH; by decompn. of the Na salt with HCl a yield of 30 g. was obtained from 2.2 kg. leaves.

On fractional crystn. from 75% EtOH, 4 out of 5 fractions gave a const. m. p. at 284-5° and formula $C_{11}H_{14}O_4$. *Prunol*, prepd. from leaves of *Prunus serotina* (Howe and Moore, *C. A.* 4, 2182) in the same way, and *malol*, prepd. from apple peels as previously described (Sando, *C. A.* 18, 1003) both had the same formula and m. 284-5°. The *malol* was obtained by the above method and also by fractional crystn. without conversion to the Na salt, to avoid the possibility of change by the alkali. The identity of these 3 compds., as claimed by van der Haar (*C. A.* 18, 2511), is confirmed by analysis and by crystal data, but the formula, based on 92 combustions, is as above and not $C_{11}H_{14}O_4$, as van der Haar states. The *di Ac derivs.* of all 3 compds., prepd. in good yield by boiling several hrs. with excess Ac_2O , were also identical. From 4.56 g. of the *di Ac deriv.* by refluxing 3 hrs. with 70% EtOH, 4 g. of the *mono Ac deriv.*, $C_{11}H_{14}(OAc)CO_2H$ (II), m. 289-90°, was obtained. From I was prepd. by boiling with excess Me_2SO , in MeOH and NaOH, extn. with Et_2O and crysto. from 70% EtOH, *Me ursolate* (III), $HOC_{11}H_{14}CO_2Me$ m. 170.5-15°. Acetylation of III gave a *mono Ac deriv.*, m. 246-7°. Prepn. of the same compd. by action of $SOCl_2$ on II, followed by boiling with MeOH, confirmed the *HO acid* formula. The name *ursolic acid* is therefore adopted for I, for *prunol* and for *malol*. The following *derivs.* were also prepd.: from 1 g. I by treatment with excess $\alpha-C_6H_4(CO)_2O$, 1 g. *phthalylursolic acid*, m. 264-5°, from 1 g. III by similar treatment, 0.7 g. *Me phthalylursolate*, m. 214-5°, from 2.3 g. I, suspended in 70% aq. EtOH, by making nearly alk. and boiling 1.5 hrs. with 1 g. $BzCl$, 2.85 g. *phenacylursolate*, m. 199-200°. K. V. THIMANN

Alkybenzochwood lignins and their cleavage (VON WACKER) 23. The porphyrins (FISCHER, *et al.*) 11A. Electrolytic reduction of 4 keto-3 phenyl-3,4-dihydroquinazoline (STOMI) 4. The alleged electrochemical sulfonation of an aromatic hydrocarbon (FISCHER, *et al.*) 4. New methods of organic thermochemistry (REBER) 2. Study of E. Fischer's polypeptides with x rays (LEVEL) 2. The reduction potential of thymoquinhydrone (BILMANN, MUESS) 2. Electric moment and molecular structure (SWYNN, WALLS) 2. The work of the International Bureau of Physical-Chemical Standards. IV. Study of the physical constants of 20 organic compounds (TIMMERMAN, HERNANDEZ ROLAND) 2. The isolation of the isomers of heptane from petroleum (BRUNN, HICKS-BRUNN) 22. Determination of C in sewage and industrial wastes (MOHLMAN, EDWARDS) 14. Chlorophyll formation (NOACK, KIESLING) 11D. Fe oxide pigment (U. S. pat. 1,793,941) 26

Oxygenated organic compounds by synthesis from hydrogen and oxides of carbon. H. DREYFUS. Brit. 337,409, July 24, 1929. Compds. such as EtOH, acetaldehyde, HOAc and their homologs and esters are prepd. from H and C oxides or from compds. such as MeOH synthesized from them (with or without presence also of H and C oxides) by use of catalysts such as Fe, Ni or Co borates, silicates, phosphates or salts of other oxyacids of P. Various details and examples are given, and app. may be used having Cu or Cu lined chambers or formed of steel contg. Mo, Mn, Co or Ni. Temps. up to 600° and pressures up to 500 atm. are mentioned as suitable. Cf. *C. A.* 25, 1839.

Oxygenated organic compounds from methane and steam. H. DREYFUS. Brit. 337,410, July 24, 1929. Compds. such as EtOH, HOAc, AcCl and their homologs and ketones are formed from CH_4 and steam at 200-500° and pressures up to 500 atm. or more, and the products may be varied by adding to the reaction mixt. C oxides, H or O, and inert gases such as N also may be present. Among the catalysts which may be used are Zn, Mg, Ca, Al, Cr, Mn, V, Mo, Ti, Fe, Co, Ni and rare earths, or their compds., such as oxides, As or S compds., phosphates, silicates or borates. Also there may be used mixts. of these with alkali or alk. earth compds., with each other, with Cu, Ag or Au, with compds. such as Zn vanadate, Zn chromate, Mn chromate, Zn tungstate or molybdate or with the corresponding salts of Mg, Ca or Al. App. of Cu or lined with Cu or of steel contg. V, Mn, Ni or Co is suitable.

Olefinic or aromatic hydrocarbons by catalytic conversion. I. G. FARBEYND. A.-G. Brit. 337,434, May 30, 1929. Chambers serving for conversions of this character are alternately heated (suitably to about 900°) and initial materials intermittently passed through them, while in assoc. operation with heat regenerators, and the heat regenerators are preferably made of material catalyzing the desired reactions such as firebricks impregnated with Cu salts or bricks made from fireclay and Si, or other refractory materials may be used such as alloys of Cr and Ni, ferro-Si or SiC. Initial materials which may be used include gases obtained by the degasification of carbonaceous materials with steam and (or) gases contg. O, and reaction gases from the de-

structive hydrogenation of coals, tars, mineral oils, etc., contg. homologs of CH_4 . Various details and examples of procedure are given.

▲ Olefin oxides from chlorohydrins. JOSEF WERTZ, HANS SCHRADER and ERICH WERNERBUCK (to Th. Goldschmidt A.-G.). U. S. 1,702,608, Feb. 17. In reactions such as the production of ethylene and propylene oxides from chlorohydrins, there is established and maintained (in an app. which is described) a slowly flowing body of boiling aq. liquid having an extensive surface, and to this, at a point near the beginning of flow, there is supplied corresponding quantities of chlorohydrin soln. and of alkali, and the condensed olefin oxide vapors are collected.

Organic isocollolds. L. AUER and L. SUSZTEK. Brit. 337,735, April 21, 1920. See 1r 601,310 (C. A. 25, 1839).

Aminoalkyl and alkylaminoalkyl compounds. JOSEPH KLAER and FRITZ MERTSCH (to Winthrop Chemical Co.). U. S. 1,703,170, Feb. 17. A described process comprises the reaction of a compound of the general formula $\text{R}_1\text{R}_2\text{N}$ alkylene halogen, wherein R_1 and R_2 stand for substituents of the group consisting of H and alkyl, with a compound of the general formula halogen MgR_3 , wherein R_3 stands for a substituent of the group consisting of aryl and hydroaromatic residues and heterocyclic residues contg. N and decomposing the reaction product with water. Details are given of the production of β -1,4- $\text{N}(\text{C}_2\text{H}_5)_2$ -Ph, 1,4- $\text{N}(\text{CH}_3)_2$ -Ph, 1,6- $\text{N}(\text{CH}_3)_2$ -Ph, 1,6- $\text{N}(\text{CH}_2\text{CH}_2\text{C}_6\text{H}_4)_2\text{OMe}$ -4, 1-Et- $\text{NCH}_2\text{CH}_2\text{C}_6\text{H}_4$ -1r, 2-methyl- β -diethylaminoethylindole (b.p. 171° and forming white crystals after manipulation with ether), 1-Me- $\text{NCH}_2\text{CH}_2\text{CH}_2\text{C}_6\text{H}_4\text{OMe}$ -4 and a base of the formula $\text{PhCH}_2\text{CH}_2\text{NH}_2$.

Replacing amino groups by halogen in aromatic and aliphatic compounds. RICH. AAR KJØM and ERNST LICHTENBERGER. Swiss 141,750, Oct. 30, 1928. Aided to 136,900 (C. A. 24, 1304). The method of 136,900 is improved by removing the N atom from the diazonium halide compd., formed by diazotizing the amine in the presence of a high halogen concn., by oxidation. In the example, the diazonium chloride obtained by diazotizing *o*-toluidine in the presence of a paste of ZnCl_2 and NH_4Cl , is subjected to the radiation of a Hg lamp, this induces the evolution of N.

Methyarylaminos. EDGAR C. BRITTON and WM. H. WILLIAMS (to Dow Chemical Co.). U. S. 1,703,093, Feb. 21. In making an *N*-methyl deriv. of an arylamine such as dimethyl *p*-toluidine, reaction is effected between the arylamine and MeOH with the aid of a relatively small proportion of MeBr.

Sulfonating aromatic amines. ARTHUR R. MURPHY and JOSEPH B. ONSEU (to Newport Chemical Corp.). U. S. 1,701,561, March 3. For introducing a sulfo group in *o*-position to the NH_2 group, an aromatic amino compd. such as *m*-xylydine or β -naphthylamine is treated with SO_3 in an inert solvent such as $\text{C}_6\text{H}_5\text{Cl}$.

Polymers of vinyl alcohol. I. G. FARNENIND A.-G. (Fritz Klatte and Arthur Zimmermann, inventors). Ger. 514,593, Nov. 13, 1928. Halogenated polyvinyl esters are saponified with NH_4OH . Thus, polyvinyl chloroacetate is saponified with aq. NH_4OH .

Esters of organic acids and polyhydric alcohols. I. G. FARNENIND A.-G. (Martin Luther and Walter von Kmling, inventors). Ger. 514,503, June 10, 1927. Small quantities of metal soaps are added during the esterification in the liquid phase. Thus, stochiometric amts. of olive oil, fatty acid and glycerol are heated in a high vacuum to 170–250° with addn. of 0.2% Mg oleate.

Esters of glycolic acid. I. G. FARNENIND A.-G. Brit. 337,609, Nov. 8, 1920. Dry anhyd. alkali metal or alk. earth metal salts of chloroacetic acid are heated with aliphatic alcs. such as ethyl, butyl or allyl alcs., 1,3-butyleneglycol, glycerol or cyclohexanol, in the presence of an esterifying agent (generally above 100° and suitably under pressure). Examples are given of the production of the ethyl, butyl and allyl esters and that of 1,3-butyleneglycol.

Esters and amides of 2-aminonaphthalene-3-carboxylic acid. I. G. FARNENIND A.-G. (Ludwig Sander, inventor). Ger. 514,590, July 15, 1928. Alc. phenol, NH_4OH , or primary or secondary amines (except acyclic aliphatic secondary amines) is caused to react on 2,3-naphthoic acid anhydride. Bases may be present. Thus, alc. NH_4OH contg. 5% NH_3 , if stirred with the anhydride for 3 hrs. at 50°, gives a 60% yield of 2-aminonaphthalene-1-carboxamide, m. 231–236°. Several other examples are given.

Stabilizing thioacarbamide solutions. I. G. FARNENIND A.-G. (Philipp Siedler and August Moeller, inventors). Ger. 494,910, Jan. 20, 1928. Supersatd. solns. of thioacarbamide in aromatic amines or heterocyclic bases are stabilized by addn. of synthetic or natural resin. Thus, a soln. of thioacarbamide in *o*-toluidine is stabilized by addn. of colophony, turpentine basic waste resin and thick tar and cooling to –15° for 2 hrs.

Purifying benzanthrone. IVAN GURELMAN, ROBERT J. GOODRICH and EDWARD T. HOWELL (to Newport Chemical Corp.) U. S. 1,794,545, March 3. A reaction mass comprising benzanthrone and assorted impurities in concd. H_2SO_4 is dild. with water to a concn. of H_2SO_4 in which benzanthrone is relatively insol., and the benzanthrone is extrd. by means of a selective solvent such as toluene which is immiscible with the H_2SO_4 and is then recovered from the solvent. Cf. C. A. 25, 1539.

Substituted guanidines. EDWARD TSCHUNKER and HELMUTH MEIS (to I. G. Farbenind. A.-G.). U. S. 1,794,875, March 3. See Ger. 592,045 (C. A. 24, 4791).

Acenaphthene derivatives. J. G. FARRENING A.-G. (Ernst Runne, inventor). Ger. 517,554, Aug. 27, 1929. 5-Hydroxyacenaphthene and its derivs. are prepd. by heating 5-aminonaphthalene and its derivs. with water and a mineral acid or acid salt. Thus, 5-aminonaphthalene may be heated for 8 hrs. at 200° with H_2SO_4 of about 10% concn. The prepn. of 5,6-dihydroxyacenaphthene, m. 195° , from the diamino compd. is also described.

Benzophenone derivatives. I. G. FARRENING A.-G. (Sebastian Gasser and Bernhard Ewert, inventors). Ger. 517,478, Feb. 14, 1929. Di- and tri-carboxylic acids of alkoxybenzophenones and their substitution products are prepd. by oxidizing alkoxy mono- or di-methylbenzoyl-o-benzene acids or their substitution products with KMnO_4 in alk. soln. Thus, 4'-methoxy-5'-methylbenzoyl-o-benzene acid yields 4'-methoxybenzophenone-2,5'-dicarboxylic acid, m. $234-5^\circ$. Examples are given also of the prepn. of 2'-methoxybenzophenone-2,4'-dicarboxylic acid, m. $247-50^\circ$, 2'-methoxybenzophenone-2,5'-dicarboxylic acid, m. $238-9^\circ$, 2'-methoxy-5'-chlorobenzophenone-2,4'-dicarboxylic acid, m. $213-4^\circ$, 2'-methoxybenzophenone-2,5'-dicarboxylic acid, m. $140-2^\circ$, 2'-methoxy-3'-chlorobenzophenone-2,5'-dicarboxylic acid, m. $234-6^\circ$, and 2'-methoxy-3,5'-dichlorobenzophenone-2,4'-dicarboxylic acid, m. $219-21^\circ$. Cf. C. A. 25, 1538.

Safrole derivatives such as protocatechuic aldehyde and isoeugenol. FRITS E. STOCKMACH. U. S. 1,792,716, Feb. 17. It has been found that, contrary to previous investigations, safrole yields, when heated with alc. potash to a temp. of $140-170^\circ$, a mixt. of two isomers, namely, 4-allylmethoxy-3-hydroxy-1-propenylbenzene, and 4-hydroxy-3-allylmethoxy-1-propenylbenzene, and that when this mixt. is subjected to the action of certain oxidizing agents, only the 4-hydroxy isomer is oxidized, while the 3-hydroxy isomer remains unoxidized and is readily recoverable. This makes possible the com. production of protocatechuic aldehyde and of obtaining as a by-product isoeugenol. Numerous details of procedure are described.

Safrole derivatives such as vanillin and isochavicol. FRITS E. STOCKMACH. U. S. 1,792,717, Feb. 17. When the isomeric reaction product produced by heating safrole with alc. KOH at $140-170^\circ$ (see preceding abstr.) is methylated a product containing a mixt. of methyl ethers of the two isomers is obtained, and upon heating this resulting product in a weak acid soln., the mixed ether group is split off with the evolution of formaldehyde and an alc. and the formation of isoeugenol and its isomer isochavicol. When this mixt. of isoeugenol and isochavicol is subjected to the action of certain oxidizing agents, only the 4-hydroxy isomer, that is, isoeugenol, is oxidized, while the 3-hydroxy isomer, that is, isochavicol, remains unoxidized and is readily recoverable. This makes possible the com. production of vanillin from safrole, and of obtaining as a by-product isochavicol. Various details of procedure are described.

N-Substituted derivatives of aromatic aminohydroxy and polyamino compounds. I. G. FARRENING A.-G. Ger. 514,757, June 18, 1927. Addn. to 499,825. The methods of 499,825 (C. A. 24, 4321) and 512,456 (C. A. 25, 1629) for producing N-substituted amines of aromatic aminohydroxy and polyamino compds. are modified by replacing the aliphatic, heterocyclic and hydroaromatic linked N, by N in the form of alkylammonium compds. containing two or more N atoms capable of conversion into strongly basic polyamino compds. Thus, 1-amino-3-methoxy-4-isopropoxybenzene and the dihydrochloride of ethyldiethylammonioethylammonioethyl chloride are fused together at $100-110^\circ$ for 8 hrs. to give the base 3-MeO-4-iso-PrOC $_2\text{H}_5\text{NHCH}_2\text{CH}_2\text{N}^+\text{EtCH}_2\text{CH}_2\text{N}^+\text{Et}$. Other examples are given.

Organic compounds of lithium. KARL ZINGLER. (Herbert Colomus, inventor). Ger. 512,852, Oct. 15, 1929. Org. compds. of Li with the metal linked directly to C are obtained by the reaction of org. halogen compds. with Li in a solvent indifferent to Li alkyl and Li aryl compds., at temps. at which the reaction velocity between the Li alkyl or aryl compds. and the halogen alkyl or aryl compds. is low. Thus, Li wire or dust is added to abs. ether and PhBr added to produce a 70-80% yield of PhLi and LiBr . Other examples describe the production of MeLi and BuLi.

Aliphatic arseno compounds. CHARLES S. PALMER (to Parke, Davis & Co.).

of CO_2 on the Na salt of thiophenol nt $150-190^\circ$ and 30-50 atm. Other examples are given

1-Methyl-5-chlorobenzene-2-carboxamide-3-thioglycolic acid. J. G. FARBENING A-G (Norbert Steiger, Erwin Hoffa and Hans Ilcyna, inventors). Ger 514,505, April 1, 1927 1-Methyl 2-cyano 5-chlorobenzene 3 sulfonic acid chloride is reduced with a metallic reducing agent in the presence of a strong mineral or org acid and with addn of an inert org solvent for the sulfonic acid chloride, and condensing the resulting 1-methyl-5-chlorobenzene 2-carboxamide 3 mercaptan with $\text{ClCH}_2\text{CO}_2\text{H}$ at suitable temps Examples are given

Purifying sulfonic acids GAIGORI PETKOFF Ger 517,150, Jan 18, 1927. Sulfonic acids of high mol wt are freed from unsulfonated material, resins, H_2SO_4 , etc., by mixing them, or their solns in water and (or) alc, with porous cellulosic material, e g, saw dust, wood meal or hydrocellulose The mixt is dried and then extd. in turn with benzine, C_6H_6 , and LiOH The first extn removes mineral oils, resins, etc., the sulfonic acids being taken up in the later extns H_2SO_4 remains in the cellulosic material. The method may be applied to octahydroanthracenesulfonic acid, alkyl-naphthalenesulfonic acids, sulfo aromatic fatty acids, etc. An example is given.

Anhydro-*N*-pyridiniumsulfonic acid. PAUL BAUMGARTEN Ger 514,821, Mar 17, 1926 SO_2 or SO_2 -yielding agents such as oleum, ClHSO_3 or its esters, or SO_2Cl_2 is caused to react on pyridine or its *C*-alkyl homologs, alone or in presence of a diluent. Thus, a 90% yield is obtained by the action of SO_2 on a soln. of pyridine in CCl_4 . Other examples are given

Acetic anhydride, etc. H. DREYFUS. Brit 330,668, July 10, 1929 Aliphatic acid anhydrides are made by heating the vapor of the corresponding acid in contact with a catalyst supported upon or mixed with asbestos, e g, HIOAc vapor is passed rapidly through a Cu tube at $400-500^\circ$ contg Ca tungstate upon asbestos and the anhydride is formed by fractional condensation Cf C A 25, 1538.

Tetrahydrofurfuryl alcohol. GEORGE D. GRAVES (to E. I. du Pont de Nemours & Co) U S 1,794,453, March 3 Furfural is heated under pressure with H, a Ni catalyst and water

Formaldehyde from oxidation of hydrocarbons. H. HARTER Brit. 337,407, June 24, 1929 An app is described for oxidizing gaseous or readily volatile hydrocarbons by air (or air enriched with O) which, just before use in the reaction, is passed through an elec. flaming arc by which N oxides are formed which facilitate the production of CH_2O . The oxidation may be assisted by catalysts such as Cu, Ag, Co or Mn, assocd with silica gel or active C

Acetaldehyde from acetylene. J. G. FARBENING A-G Ger. 514,501, Nov 9, 1918 Addn to 504,862 (C A, 25, 524) C_2H_2 is treated with catalysts consisting of Fe compds which are converted by C_2H_2 into reducible oxides Thus, powd. hydrated Na_2SiO_3 , $\text{Al}_2(\text{SiO}_3)_3$ is suspended in a soln of FeCl_3 and the Fe pptd by concd NH_4OH . The ppt is washed, pressed and dried, and C_2H_2 is led over it at 350° to give AcH . Other examples are given

Acetone from isopropyl alcohol. N-V. DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ Brit 337,666, Oct. 10, 1929 Dehydrogenation of isopropyl alc. at temps below 400° (preferably not over 350°) is effected in the presence of catalysts such as Cu, MnO_2 , BaO , metals of the Pt group, ZnO , CdO , blue W oxide, Mn oxide, V pentoxide, MgO , or oxides of Be or Zr, with the addn of small quantities (preferably not over 1%) of Th oxide, Na_2CO_3 , Ce oxide, or of oxides of Zn or Zr if these are used merely as auxiliary addns The catalyst may be regenerated with steam or air or both, by heating followed by reduction

Urea. J. G. FARBENING A-G Brit. 337,394, July 29, 1929 Various details are described relating to the distn. of crude melts contg urea together with NH_3 , carbonate or NH_4 carbamate and water, obtained in the production of urea from NH_3 and CO_2 under pressure, to recover NH_3 and CO_2 . During the stage of the distn effected at temps over 150° under pressures sufficient to cause the distillate to collect as a liquid, an excess of NH_3 , over that required to combine with all the CO_2 present in the distn. app is continuously maintained in the app in order to check decompn of urea

Nitrobenzene purification. SYLVESTER BOYER (to General Elec Co). U. S. 1,793,304, Feb 17. The material is passed through fuller's earth, then treated with Al oxide and distd

1,2,3-Trihalobenzenes. J. G. FARBENING A-G (Richard Herz and Eduard Albrecht, inventors). Ger 517,429, Oct 6, 1927 See Fr 661,490 (C A, 24, 379).

Dimethylaniline. EDGAR C. BRITTON and WM. H. WILLIAMS (to Dow Chemical

Co.). U. S. 1,794,057, Feb. 24. A mixt. of aniline and MeOH is heated with addn. of a small proportion of MeBr.

2,6-Xylidine. IMPERIAL CHEMICAL INDUSTRIES, LTD. Ger. 517,338, Mar. 1, 1930. See Brit. 328,418 (C. A. 24, 5309).

Benzidine and its substitution products. HENRY J. WEILAND and IVAN GUBELMANN (to Newport Chemical Corp.) U. S. 1,791,097, Feb. 24. For effecting purification of benzidine or derivs. such as *o*-tolidine or *bi-o*-anisidine, a salt of the impure material such as the hydrochloride or sulfate is treated with an alk. aq. soln. such as NaOH to give an alk. reaction to the resulting soln. and with a sufficient quantity of a water-immiscible org. solvent to dissolve the benzidine compd. when hot, insol. impurities are sepd., the benzidine compd. is crystd. from the org. solvent and is recovered without any intermediate isolation.

Carbazoles. I. G. FARBENIND A. G. (Friedrich Stolz, Walter Kross, Gustav Ehrhart and Hans Schliehenmaier, inventors) Ger. 514,822, Jan. 29, 1929. Perhydrocoated carbazoles are obtained by treating carbazole or its N-alkyl derivs. with H₂ under pressure and at high temps. in the presence of a catalyzer formed by heating a Ni salt alone or mixed with another metal (preferably of the Fe group) with a carrier such as SiO₂, to a temp. of above 400°. Examples are given. Cf. C. A. 25, 801.

Isatins. I. G. FARBENIND A. G. (Hans Colombara, inventor) Ger. 514,595, Mar. 28, 1928. Isatins are obtained by treating oxamyl halides of primary amines of the general formula R-NH-CO-CO-halogen (R = a residue of the benzene series, with at least one substituent) with a condensing agent. Thus, 5,7-dimethylisatin (m. 242°) is obtained by treating 2,4-dimethyl-1-phenyloxamyl chloride with AlCl₃. Many other examples are given.

1-Phenyl-2,3,4-trimethyl-5-pyrazolone. I. G. FARBENIND A. G. (Max Bockmühl, Karl Streitwolf, Alfred Fehle and Walter Herrmann, inventors) Ger. 514,823, May 7, 1929. More than 6 mols. of CH₃O is caused to react on 1-phenyl-2,3-dimethyl-5-pyrazolone or on methylenebis(1-phenyl-2,3-dimethyl-5-pyrazolone) in the presence of concd. HCl. The reaction takes place at raised temp. and under pressure. Examples are given. Cf. C. A. 24, 4521.

Antipyrine. ERNEST F. GREYER and JOSEPH P. BELSLEY (to Dow Chemical Co.), U. S. 1,792,833, Feb. 17. Methylation of compds. such as 1-phenyl-3-methyl-5-pyrazolone in the production of antipyrine is effected with MeOH and an alkyl halide such as EtBr, the alkyl radical of which contains at least 2 C atoms.

α -Hydroxy- β' -iodopyridine. CURT RATH. U. S. 1,793,962, Feb. 24. In forming compds. of this character, an α -substituted β' -aminopyridine is diazotized and the diazo soln. is caused to react with a metal iodide such as KI. The α -hydroxy- β' -iodopyridine is a white cryst. compd. sol. in water. Cf. C. A. 24, 2472.

Menthol. WALTER SCHOELLER, HANS JORDAN and REINHARD CLERC (to Schering-Kahlbaum A. G.), U. S. 1,793,020, Feb. 17. See Ger. 508,995 (C. A. 25, 717).

Menthol. SCHERING-KAHLBAUM A. G. (Walter Schoeller, Hans Jordan and Reinhard Clerc, inventors) Ger. 514,594, July 16, 1927. Addn. to 512,719 (C. A. 25, 1260). The condensation product from acetone and *m*-cresol is acetylated and the product hydrogenated until 8 H atoms are taken up. Sapon. then gives menthol. Examples are given. Cf. C. A. 25, 717.

Sterols. I. G. FARBENIND A. G. (Robert Griessbach and Otto Ambros, inventors). Ger. 517,499, April 7, 1927. See Brit. 322,465 (C. A. 24, 2760).

Styrene and its homologs. I. G. FARBENIND A. G. Brit. 338,262, Sept. 10, 1929. See Fr. 682,569 (C. A. 24, 4527).

11—BIOLOGICAL CHEMISTRY

PAUL E. HOWE

A—GENERAL

FRANK P. UNDERHILL

The birth of chemical biology. J. B. LEATHES. *Brit. Med. J.* 1930, II, 671-6.—A lecture.

J. B. BROWN

The porphyrins. XXVII. The mechanism of iron introduction into porphyrins and isolation of crystalline hemes. H. FISCHER, ALFRED TREIBS and KARL ZEILE. *Z. physiol. Chem.* 195, 1-27 (1931), cf. C. A. 25, 118.—In the various hemins the Fe is

tervalent. With few exceptions all contain the group FeCl substituted in the NH groups of 2 pyrrole rings. A curious fact is that the prepn of these Fe complexes from porphyrins is best accomplished by the use of Fe^{++} instead of Fe^{+++} . The explanation of Hauserowitz that a (N)Ic complex is first formed which immediately reacts with HCl to form (N)- FeCl and H was tested experimentally with protoporphyrin in an atm. of N. Neither free H nor the hydrogenation product mesoporphyrin could be detected, the product obtained was hematoporphyrin. On the other hand, the treatment of mesoporphyrin with Fe^{++} in the absence of O gave a product with the characteristic spectrum of hemin. The reaction is strikingly demonstrated by the use of an inverted V-tube, one arm contg a soln of porphyrin in AcOH , a little HCl and $\text{Fe}(\text{OAc})_3$. The air is displaced by N and the app sealed. When heat is applied to the mixt. the HCl distills to the other arm and the residual soln suddenly changes color because of the formation of Fe complex. If the distillate is now poured back by tilting the tube the reaction is reversed and the original color restored. The process may be repeated 5-10 times until finally a partial oxidation results from secondary reactions. After air has been admitted the complex formation is no longer reversible. With FeCl_3 and pure AcOH or PrCO_2H the complex is not formed unless high temps. are used, and then only partially, because of a reduction of Fe^{+++} to Fe^{++} by decomposition products. The addn of Fe to a porphyrin thus occurs via an Fe^{++} complex. Such complexes for which the term *heme* is proposed (cf Anson and Mirsky *C A* 23, 1428) have now been prepd and isolated in cryst. form from etio-, proto- and meso-porphyrin and its ester. All are Cl-free even in the presence of Cl ions. When treated with pyridine they give an intense hemochromogen spectrum. The hemes are extremely unstable to HCl , the Fe is split off as Fe^{++} as shown by the Turnbull blue reaction with $\text{K}_4\text{Fe}(\text{CN})_6$. They are easily oxidized to the stable Fe^{+++} complex especially in the presence of Cl ions. Only Fe^{++} can be introduced into the porphyrins. Since the Fe^{++} complex is more easily oxidized than the Fe^{+++} salt, only a min. quantity of the latter is required, because the Fe^{+++} salt present effects the oxidation of the Fe^{++} complex. This institutes a chain reaction which proceeds almost as rapidly as that in which the required quantity of Fe^{++} is initially present. Atm. O plays an important role, as also the catalytic property of hemin itself. By using a technique which excludes the possibility of oxidation, *etioheme*, *protoheme*, *mesoheme* and *meso-ester heme* were obtained cryst. from the corresponding porphyrins. In the dry state they are more stable to atm. O than their solns. A cryst. hemochromogen was prepd. by addn of pyridine to the meso-ester heme. Hemoglobin is probably a mol. compd. of heme with globin, the latter effecting a stabilization of the mol. O converts it to oxyhemoglobin which is still more stable. Methemoglobin, on the other hand, contains hematin with Fe^{+++} and is extremely stable to reducing agents. It is surprising that the spectrum of hemoglobin bears no resemblance to that of hemochromogen, such as might be expected on the assumption of a mol. compd. of heme and globin. A possible explanation may be its peculiar state of dispersion, since reduced hematin with native globin gives hemochromogen.

A. W. DOX

The effect of radioactive emanations on the automatic activity of the frog heart with particular reference to the action of large doses of such radiations and to the antagonism of α - and β -rays. E. POLAK. *Arch ges Physiol (Pflügers)* 225, 653-75(1931).

ARTHUR GROLLMAN

Purinoxytic enzymes of the human organism. RICHARD TRUSKOWSKI. *Biochem J* 24, 1681-5(1930).—Uricase was found to be absent from the livers of a new born child and in 2 adults. Xanthine oxidase is present in liver exts. from the 12th to the 40th week of pregnancy.

BENJAMIN HARROW

Denaturation of proteins. VII. Denaturation in the presence of alcohol. NORMAN BOOTH. *Biochem J* 24, 1699-1705(1930), cf. *C A* 24, 5314.—The speed of denaturation of hemoglobin has been detd. in the presence of alc. ranging from 0 to 30 vols. %, the temp. varying from 20° to 60.5°, and the pH from 5.19 to 7.74. The min. point on the pH velocity curve was shifted considerably from the neutrality point of water to the acid side by the addn. of alc. The critical increment of activation increased appreciably with the amt. of alc. present.

BENJAMIN HARROW

Effect of light and salts on gelatin. ADA GALINSKY. *Biochem J* 24, 1706-15(1930).—When gelatin is made insol by treatment with dichromate and exposed to light, a change in the protein similar to denaturation takes place, but this involves no change in the Hausmann nos., nor in the digestibility of the protein. Cr is fixed as its sesquioxide. Coagulation probably takes place in 2 stages, the first being a chem. action, and the second, which is phys. in nature, produces the coagulation.

B. H.

Enzymes and vitamins in present-day chemistry. H. C. SHERMAN. *J. Chem Education* 8, 652-60(1931)

E. H.

Physico-chemical studies of phosphatides. KORETOSHI OSE. Gifu Imp Coll. of Agr (Japan), *Research Bull* 2, 30 pp (1930) (in German).—Com egg yolk lecithin electrolyzed for 1 month resulted in isolation of 2 cryst substances, one white, m 57.5°, $C_{11}H_{19}N_3P_2O_{11}$, the other brown, m 50.5°, $C_{11}H_{19}N_3P_2O_{11}$. The latter is readily sol in cold 95% EtOH, the former with difficulty. K. KITSUTA

A case of sterilization in man due to Röntgen rays. H. MADISSON. *Deut. Z. ges. gerischil Med.* 16, 210-5 (1931). FRANCES KRASNOW

The story of histamine. D. ACKERMANN. *Z. Biol* 91, 73-4 (1931). FRANCES KRASNOW

The disintegration of proteins by amides. Disintegration of casein in acetamide. E. CIPRANULIEZ and G. DE MANOROT. *Helv. Chim. Acta* 14, 163-83 (1931).—Proteins in general are sol in acid amides and when such solns are heated, the proteins are disintegrated yielding substances of relatively low mol wt. For instance, when finely divided casein suspended in molten AcNH₂ is heated to 200°, a mixt is obtained, the mol wt of which fluctuates between 217 and 390. The elementary analysis and color reactions (biuret, Millon's, xanthoproteic) of the mixt were similar to those of casein. However, the COOH groups were free but the NH₂ groups were not. When the casein was subjected to the heat treatment 5 min or less, the products gave a pos biuret test immediately, otherwise the test was obtained only when the mixt was allowed to remain in the presence of alkali for some time. C and M consider that the long heat treatment caused a profound change in the casein mol and that the alkali polymerized the products to compds capable of yielding the biuret test. They agree with previous workers that the open chain mols in casein are broken and transformed into closed chain substances of lower mol wt. I. M. LIVING

Denaturation of proteins by urea and related substances. F. GOWLAND HOPKINS. *Nature* 126, 328-30, 343-4 (1930).—If in a few cc of egg albumin recrystd by the method of Hopkins and Pinkus and freed from the NH₄ sulfate, or in a soln of ordinary egg white, urea be dissolved and the soln evaporated in vacuum, the urea denaturizes the protein so that the former can be extd with H₂O, while the protein becomes wholly insol. A change in the pH from 4 to 7 does not affect the acquired condition of the protein. The denaturated protein gives a red-purple color characteristic of a sulfhydryl reaction when tested with nitroprusside and NH₄OH. Nondenaturated proteins do not give this reaction, indicating that a mol rearrangement has taken place. The assumption is made that the precursor of the active thiol group is a disulfide grouping not present in the native protein, but established on denaturation. Pos results obtained with certain N-contg substances and neg with certain others indicated that to some degree at least a relationship exists between denaturing power and constitution. An amide structure appears to be necessary, and in dialkyl compds 1 amino group apparently must remain active. The effect of urea on protein varies in detail with the abs and relative concns of the 2 constituents. Concomitant with the mol rearrangement are profound changes in the colloidal and general phys properties of the protein. A study of these resulted in the development of a simple method of detg the degree and rate of protein denaturation, which in its turn was used for further study of the phenomenon of protein denaturation. Serum proteins, on the whole, are affected by similar denaturants in a similar manner on evapn. Observed in soln their behavior indicated that blood proteins are more resistant to dehydration, which characterizes the change from lyophil to the lyophobic condition. The behavior of euglobins has not been studied. B. S. LIVING

The absence of asparagine among the biuret-free products of proteins hydrolyzed by enzymes. A. CLEMENTI and D. TORRESI. *Boll. soc. ital. biol. sper.* 5, 956-8 (1930).—During the digestion of casein, egg albumin and peptone to the point of their complete transformation into biuret-free products by the successive action of proteolytic enzymes present in the gastric, pancreatic and enteric juices of the dog, and thus without the intervention of asparaginase, 95% (often 100%) of the amide N is set free as NH₃. The amide N present in biuret-free products is set free completely as NH₃ by guinea-pig serum in which asparaginase is present, by dog serum, or by dog enteric juice in which asparaginase is surely absent. Conclusion: During digestion of the protein mol aspartic acid is liberated as such and not in the form of asparagine. Asparagine cannot be considered as one of the intermediate enzyme degradation products of protein. PETER MASUCCI

The action of phosphatase of the bones on glycerophosphoric acid. A. DE RRENZO. *Boll. soc. ital. biol. sper.* 5, 1031-6 (1930).—The action of phosphatase on asymmetric α -glycerophosphoric acid was studied. The enzyme was prepd according to the Robison method and purified by electroultrafiltration. The α -glycerophosphoric acid was prepd according to the Karrer method. Five cc. of Na α -glycerophosphate (0.1332% soln),

2 cc. of borate buffer soln. of varying pH values and 0.5 cc. of a 1.5% soln. of the enzyme were placed in each of 3 test tubes. In another set, an equal vol. of equimol. soln. of β -glycerophosphate was used. After 24 hrs. at 37° the morg. P in each tube was detd. by the Bell and Douy colorimetric method. The results obtained show that bone phosphatase hydrolyzes α -glycerophosphoric acid more readily than β glycerophosphoric acid. PETER MACVICK

The kinetics of fermentation processes. The kinetics of invertase. E. AVTONOMOVA. *Biochem. Z.* 231, 13-24 (1931).—The invertase reaction like the catalase reaction may follow the equation of a monomol. reaction but under various conditions the constants, on this basis may be either increasing or decreasing. Decreasing constants for invertase are obtained when a very low sugar concn. and relatively high enzyme concn. are used. On the contrary, with high concn. of catalase and low concn. of H_2O_2 the constants are increasing. Obviously enzymes whose activity tends to deviate from the monomol. type do so in a definite direction, and in the case of catalase this is predominantly toward decreasing, whereas in the case of invertase toward increasing values. This is not due to differences in the nature of these enzymes but rather to the fact that the H_2O_2 is a powerful agent which destroys the catalase whereas the invertase is inhibited by the reaction products. The following equations are proposed to describe every instance of the enzyme kinetics where the activity either decreases or increases: $C_1 = 1/(A - D) \times \log D(A - X)/A(D - X)$ and $C_2 = 1/(A + D) \times \log A(D + X)/D(A - X)$, resp. The value D is proportional to the initial enzyme activity. S. MOROGLIS

Changes in the chloride distribution in blood under the influence of ultra-violet radiation. JERRY GLASS. *Biochem. Z.* 231, 45-53 (1931).—Intensive ultra violet radiation causes 2-4 hrs. after the treatment a change in the Cl distribution in the blood, with a tendency to accumulation in the red cells, which may persist for 24 hrs. This indicates a shift in the blood acid base equil. toward an acidotic condition. S. M. OELKEAS

Enzymic formation of an ester of mandelic acid. P. ROVA, R. AMONOV AND H. A. OELKEAS. *Biochem. Z.* 231, 59-66 (1931).—In the presence of isomolar concns. of l - and d -mandelic acid and Bu alc. the esterase from human, pig and cat liver, or from pig pancreas and kidney forms the Bu ester of mandelic acid with equal velocity. Thus, the marked specificity of these esterases which was found previously in the study of the hydrolysis of mandelic acid esters is not observed in the enzymic esterification. S. MOROGLIS

The enzymic transformation of guanidine to urea. NICOLAI N. IVANOV AND A. N. AVETISSOVA. *Biochem. Z.* 231, 67-78 (1931).—*Aspergillus niger* grown in a peptone culture can utilize guanidine very well as a source of N if this is furnished in conjunction with glucose, when the guanidine is changed to urea and NH_3 . This depends upon the presence in the mycelium of a guanidinase. The dry mycelium of *Aspergillus niger* contains only guanidinase and no urease and can change guanidine quantitatively to urea and NH_3 . Boiling this dry mycelium destroys its guanidinase activity. S. M.

The Boas benzidine reaction of the potato. T. YOSHIOKA. *Biochem. Z.* 231, 233-8 (1931).—The potato enzyme described by Boas (*C. A.* 25, 307) which reddens benzidine is entirely unaffected at 50° , but is weakened in its action at 60° and is completely inactivated at 70° . The enzyme can be pptd. by half satn. with $(NH_4)_2SO_4$ or full satn. with $MgSO_4$ and is made insol. by alc. or acetone. With kaolin one obtains an insol. ppt. which has the full enzyme activity, but the enzyme cannot be leached out. With $Al(OH)_3$ no adsorption could be obtained. KCN or $HgCl_2$ inhibits the enzyme activity. Ni powder has no effect. The enzyme shows a great tendency to become insol. Through $(NH_4)_2SO_4$ pptn. enzyme preps. can be made poor in N and of good potency, but through acetone pptn. active preps. are made which retain full enzymic activity permanently, but they are insol. S. MOROGLIS

A protein ethereal sulfate compound from spleen. ALFRED EBEL. *Biochem. Z.* 231, 306-8 (1931).—A kg. of fresh spleen is chopped up fine, mixed with 1 l. H_2O and 40 cc. N AcOH and heated to 70° . The coagulum is filtered off, and the filtrate is pptd. with 200 cc. 20% $Pb(OAc)_2$. After this is made alk. with NH_4OH the ppt. is sepd. and the clear filtrate is freed from Pb with H_2S . The final filtrate is evapd. *in vacuo* to a small vol. and again filtered. The addn. of twice the vol. of alc. causes no turbidity. The concentrate is dialyzed 48 hrs. and, after filtering, is once more concd. *in vacuo*. Treated with 10 times its vol. of alc. a thick ppt. is formed which is allowed to settle in the ice box and is collected on a hardened filter paper, washed with alc. and dried. This powder is dissolved in very little H_2O and again pptd. as before with 10 vols. alc., the ppt. washed with abs. alc. and ether. This substance gives no Millon reaction, a weak tryptophan reaction and a Molisch test for carbohydrate. The substance can be salted

out by full satn with $(\text{NiF}_2)_2\text{SO}_4$ and contains about 1.70% S which is hydrolyzable with 10% HCl at 100° S MORGULIS

Studies on phosphatase. I. The kidney phosphatase of different laboratory animals. MASAKI UMENO *Biochem Z* 231, 317-23(1931) —The phosphatase activity of the kidney from different animals has been tested on a Na glycerophosphate substrate. The relative activity follows the series: hen>cat>dog>steer>toad>rabbit>guinea pig. The phosphatase content of the various portions of the kidney was cortex>medulla>papilla. II. The liver phosphatase of different laboratory animals. *Ibid* 324-27 —The liver phosphatase activity is about half as strong as the kidney phosphatase, and is also differently distributed in the animal series: hen>rabbit>dog>guinea pig>frog. III. The phosphatase content of the kidney and liver in experimental nephritis. *Ibid* 328-33 —In exptl nephritis the kidney phosphatase activity is reduced to $1/10$ of its normal value. The liver phosphatase activity during nephritis is also diminished but only to a small degree. IV. The optimum temperature and the inactivation temperature of the kidney phosphatase. *Ibid* 334-8 —The optimum temp for the activity of the steer or dog kidney phosphatase is 42.5°. The enzyme is destroyed at a temp between 60 and 65°. V. The glycerophosphatase of leucocytes in blood. *Ibid* 339-45 —The leucocytes of a leucemia patient were found to contain much phosphatase. One cc rabbit leucocytes hydrolyzes 0.036 g Na glycerophosphate per day, while the erythrocytes show no phosphatase action. VI. The presence of phosphatase in bile and pancreatic juice. *Ibid* 346-51 —Bile contains a very strong phosphatase, whereas pancreatic juice possesses only about half as much phosphatase activity. Saliva and gastric juice are free from phosphatase, but the mucosa of the stomach, duodenum and small intestine and the salivary gland are rich in phosphatase. Phosphatase is also present in the liver, pancreas and spleen. The phosphatase content of an organ is independent of that of its secretion S MORGULIS

Resistance of insulin to certain bacteria. A. A. SCHMIDT AND KLARA TULICHINSKAJA *Biochem Z* 231, 352-64(1931) —*B. coli communis*, *Staphylococcus aureus*, *Streptococcus hemolyticus* and some anaerobic bacteria from dog feces cannot inactivate insulin. However, it is possible to obtain from feces putrefactive organisms which do destroy insulin very vigorously. It is suggested that bacteria of the peptolytic group, devoid of proteolytic enzymes, do not inactivate the insulin, which is perhaps due to the similarity between the insulin and the protein mol S MORGULIS

Influence of ptyalin on starch. II. Effect of electrolytes. J. R. BROEZE *Biochem Z* 231, 365-84(1931), cf. C. A. 23, 1917 —The action of dialyzed ptyalin sols on the starch sol in the presence of small quantities of electrolyte has been studied by changes in the expression η , $= \eta_0/\eta$, in the course of time. The symbol η , of course, refers to viscosity. Without electrolyte addn. the velocity of the change is very small, and is very much increased by the least concn. of all electrolytes, with only a small advantage in favor of cations with increasing valence. A cation lyotropic is not observed but in the univalent anions the increase in reaction velocity seems to coincide with the lyotropic series: $\text{CNS} < \text{NO}_3 < \text{Cl.F}$. The SO_4 ion does not behave according to its lyotropic position S MORGULIS

Specificity of the α -glucosidases. II. KARSTRÖM *Biochem Z* 231, 399-403 (1931) —With a strain of *Bacillus coli* which ferments maltose, glucose and fructose but not sucrose it was shown that the α -glucosidase which is operative in the hydrolysis of maltose cannot hydrolyze sucrose. S MORGULIS

The chemical constitution of serum proteins. IV. ANTON FISCHER AND ALFRED BLANKENSTEIN *Biochem Z* 231, 404-11(1931), cf. C. A. 25, 977 —Two pathological sera showed a greatly increased sedimentation reaction and an abnormally high tryptophan content. A study of the seven different chem. fractions shows that, as compared with the normal serum, there was a diminution of the albumin and an increase of the globulin content, this increase was not shared by all globulin fractions but principally by the euglobulin, which is pptd by NaCl S MORGULIS

Influence of pancreatin on collagen in the absence of neutral salts and buffer mixtures. I. A. KÖNIGEL AND O. DIETSCH *Biochem Z* 231, 423-34(1931) —The pancreatin action is exerted chiefly at the beginning of the reaction, so that prolongation of the action is safe only so far as no bacterial effect comes into play. There is practically a direct relationship between the amt. of pancreatin and the extent of the hydrolysis but the curve nevertheless shows a tendency to flatten out gradually so that an excess of pancreatin does not present any serious danger of carrying the hydrolysis too far. The enzyme activity does not increase proportionally to the rise in temp, and from about 45° the increase becomes quite remarkable, and this factor must be guarded most cautiously to prevent injury to the skins treated with pancreatin S M.

Influence of pancreatin on collagen in the absence of neutral salts and buffer mixtures. II. A. KONTTEL AND O. DIETSCHE. *Biochem. Z.* 231, 435-40 (1931). cf. preceding abstr. — This is in general a corroboration and extension of Merrill and Fleming's expts. on the digestion of collagen by pancreatin (*C. A.* 21, 1309, 2296). S. MORCILIS

Relation between swelling and proteolysis of collagen. F. NAEY. *Biochem. Z.* 231, 441-5 (1931). — Salts of benzoic, salicylic and *p*-hydroxybenzoic acids promote proteolysis in the same order as they increase the swelling. The digestion of collagen by pancreatin is definitely promoted by previous swelling. S. MORCILIS

Influence of physical factors on blood catalase. II. The influence of some physico-therapeutic procedures on the blood catalase. A. I. ALIKHEV AND K. I. RUSSKOVA. *Biochem. Z.* 231, 470-73 (1931). cf. *C. A.* 24, 1347. — The blood catalase of man under const. phys. conditions is practically the same. The outside temp. influences the blood catalase very much and should therefore be considered in the detn. of the catalase content and index. Elec. warm baths diminish the activity of the blood catalase more the higher the temp. of the bath. In a similar manner hot water baths greatly diminish the activity of the blood catalase, while a CO_2 bath or plain baths of ordinary temp. have no effect. The Scharbo baths under 2.5 atm. pressure greatly diminish the blood catalase, but this does not occur until 30 min. after the bath. Air baths produce no noticeable changes in the blood catalase content. S. MORCILIS

Specific dynamic action of protein. HENRI BORSOOK AND HOWARD M. WINEGARDEN. *Proc. Natl. Acad. Sci.* 17, 75-81 (1931). — Analysis of data on the relation of the increase in metabolism following ingestion of protein or amino acids and the efficiency of renal excretion show that the course of the sp. dynamic action parallels the course of N excretion. Of the sp. dynamic action, 25-60% is due to the work of renal excretion and the rest to N and C metabolism. This will explain previous anomalies in the sp. dynamic action of protein. The sp. dynamic action of protein is not necessarily due to the conversion of the deaminized fractions into glucose. D. S. SEARLE

The relation of life to electricity. II. The relation of stainability to electromotive force in tissues and in a variety of artificial substances, like esters, etc. R. BEUTNER AND JOS. LOHNER. *Protoplasma* 12, 52-65 (1931). — Many data are given of potential measurements of the system

+ aq. saline	ester or alc. + fatty acid	ester or alc. + amide	aq. — saline
-----------------	----------------------------------	-----------------------------	-----------------

The relation between stainability and $e. m. f.$

+ basophilic substance/acidophilic substance —

holds for fats, esters, alc., hydrocarbons etc.

B. A. SOTLE

The effect of ultra-violet radiation on the respiration of avian erythrocytes and yeast cells. GY. SZERINTI AND M. VERMES. *Magyar Orvosi Arch.* 30, 585-90 (1929). — Ultra-violet radiation under aerobic and anaerobic conditions of avian erythrocytes and yeast cells in Tyrode soln., at pH 7.2, increases the O_2 consumption, measured by the method of Warburg. The increase is about 50%. Cyanide inhibits the increased part as well as the normal respiration. This effect is reversible. H. TAUBER

X-ray analysis of bone and teeth. II. H. ROSEBERG, A. BAIRD HASTINGS AND J. K. MORSE. *J. Biol. Chem.* 90, 395-405 (1931). — From the analysis of x-ray spectrograms of bone powder a crystal structure similar to that of the apatites is deduced resembling closely the structure of dahlte, $\text{Ca}_{10}(\text{PO}_4)_6\text{Ca}_2(\text{PO}_4)_2$, in which n is not less than 2 nor more than 3. Neither CaCO_3 nor CaHPO_4 exists as such in bone. The results of de Jong (*C. A.* 20, 2602) and of Taylor and Sheard (*C. A.* 4, 24) are confirmed. $\text{Ca}_2(\text{PO}_4)_2$ also has a crystal structure similar to that of the apatites, varying slightly with different samples, but entirely different from that of CaHPO_4 . The crystal structure of enamel also resembles the apatite series, and can probably be represented as a close-packed hexagonal lattice, of which $a_0 = 20.8 \text{ \AA}$, $b_0 = 12.0 \text{ \AA}$, $c_0 = 8.82 \text{ \AA}$, referred to orthorhombic axes. The results of equilibration expts. between serum and solns. of Ca salts are considered to be of doubtful biol. significance. K. V. THIMANN

Studies on the dielectric constant of protein solutions. JEFFRIES WYMAN, JR. *J. Biol. Chem.* 90, 443-76 (1931). cf. *C. A.* 23, 5074. — The dielect. const. of solns. of rein in 70% Pr alc. was measured by the method previously described for high frequency (*C. A.* 24, 3681) and by a new resonance method at lower frequencies. At wave length 4 m. the dielect. const. of the rein solns. is lower than that of the solvent by an amt. proportional to the concn. of protein (cf. Furth, *C. A.* 17, 1579) resembling in general behavior that of

amino acid solns (cf. Hedestrand, *C. A.* 22, 3571) It rises with temp., becoming greater than the solvent at 60°. The dielec. const., however, rapidly increases with wave length up to 260 m—the longest wave used—becoming up to 50% greater than the solvent, and indicating anomalous dispersion. The zein mols. are therefore highly polar and possess a permanent elec. moment of about 60×10^{-18} e.s.u. No change in dielec. const. occurs during gelation, from which W. deduces that only a small fraction of the mols. is involved in the network, and that the viscosity of the whole gel is not the true viscosity of the soln. given by Stokes' expression, since it is to this viscosity that the relaxation times of the mols. are related. The theory of dielec. const. and the application to mol. structure are briefly discussed. K. V. THIMANN

The pH of human mixed saliva irradiation for intraoral carcinoma. GEORGE S. SHARP. *Am. J. Roentgenol. Radium Therapy* 25, 266-70(1931) —The saliva pH for the individual patient taken during the 3 phases of irradiation forms a curve, starting with an abnormally acid saliva (pH 6.1), rising during the acute reaction (pH 6.5) and returning approx. to the pre irradiation value during the post irradiation period (pH 6.0)

R. C. WILLSON

The effect of radiation on the acidity of the blood. HELEN Q. WOODARD AND HELEN R. DOWNES. *Am. J. Roentgenol. Radium Therapy* 25, 271-5(1931) —No consistent change was observed in the pH , CO_2 combining power or lactic acid content of the blood of patients after they had received one skin erythema dose of high voltage Röntgen rays. No significant change was observed in the pH of the blood of rabbits after lethal doses of high-voltage Röntgen rays

R. C. WILLSON

The effect of Röntgen rays on the inorganic phosphorus content of blood and urine of patients with carcinoma. P. I. RIVOSH AND E. R. NOVOTELNOV. *Voprosy Onkologii* 3, 150-6(1930), *J. Am. Med. Assoc.* 96, 155 —A normal P content was found in most of the 81 patients before treatment. In 2 hrs. after the treatment a steady increase in the P content of the urine and of the blood of all the patients was observed. After 24 hrs. the blood P was still increased but the urine P was normal. The same results but with larger increases in the urine and blood P were observed after Röntgen and Ra therapy of sarcoma, seminoma and lymphogranuloma. The sugar content changed in inverse proportion to the P content

R. C. WILLSON

Rotation dispersion of optically active substances (KNESEK) 2. Evolution of H_2S in the Bay of Krasnovodsk (ZABRIEV) 2. Phenomena of adsorption and protection in complex colloidal media (MARIB, MARINESCO) 2. The bile acids (WIRLAND, NOGUCHI) 10.

Peroxidase preparations. W. SAILER. *Brit.* 337,405, April 25, 1929. Vegetable substances contg. peroxidases such as embryos of rye, wheat, rice, maize or barley or parts of plants such as potatoes or brans are treated with water and fermentation is allowed to proceed in the presence of substances which in the soln. form buffer substances such as oxides, hydroxides or carbonates of alkali, alk. earths or other metals, such as ZnO , or metals not acting as catalysts. After filtering off the solids, the filtrate is treated with a preservative such as ale, benzoic acid or benzoic acid esters

B—METHODS AND APPARATUS

STANLEY K. BENEDICT

An exact rapid method for determination of water in blood and serum. M. DOLCH AND E. FÖRCHMÜLLER. *Z. physiol. Chem.* 193, 28-31(1931) —The ordinary method of H_2O detn. by drying to const. wt. does not exclude 3 possible sources of error in samples contg. protein, viz., adsorption, occlusion and chem. reaction. The method here described obviates these errors and is far more rapid, requiring only 10-15 min. It is based on the fact that the temp. at which a mixt. of $LiOH$ and xylene in definite proportions becomes cloudy varies with the H_2O content of the $LiOH$. A calibration curve is first plotted for the abs. $LiOH$ to be used. This is done by mixing 11, 11.5, 12, etc., cc. of H_2O with 100 cc. of the $LiOH$, then mixing a 20-cc. aliquot with 20 cc. of dry xylene in a suitable app. provided with thermometer and stirrer, immersing in a freezing mixt. and detg. the temp. at which turbidity develops. The curve, in which temp. is plotted against amt. of H_2O added, is a straight line. Detn. of H_2O in blood or serum is performed as follows: Pipet a 15-cc. sample into a tared 150 cc. Erlenmeyer flask and weigh the sample. Add exactly 100 cc. of the $LiOH$, mix thoroughly, boil 10 min. and allow the pptd. protein to settle. Pipet 20 cc. of the supernatant liquid into the app., add 20 cc. of xylene and det. the temp. of clouding. Locate this temp. on the calibration curve and read the corresponding wt. of H_2O from the other axis. This value $\times 100 +$ wt. of

sample is the per cent of H_2O by wt. The per cent by vol is the wt. of H_2O \times the factor 6.667. Addn of cholesterol or dried protein did not affect the results. The values obtained by this method are slightly higher than those obtained by drying the sample to const wt. and are claimed to be more accurate. A. W. DOX

A new method for the quantitative estimation of dissolved protein in gastric contents. WALTER WOLFF and FRICH JOCHMANN *Arch Verdauungskrankh* 48, 151-7 (1939). -To 5 cc of the filtered gastric contents add 5 cc of a reagent contg 10% phosphotungstic acid 0.5% HCl and 10% alc., centrifuge and wash the ppt with H_2O 3 times. Then dissolve in 3% $NaOH$, add 3 drops of 20% $CuSO_4$ soln and 3% $NaOH$ up to 10 cc. Centrifuge and compare the clear violet supernatant fluid colorimetrically with standards. Lowest values (<500 mg %) are commonly obtained in achylia gastrica. Normal values may be as high as 1000-1500 mg %. H. EAGLE

The Kaufmann method of blood-sugar determination. H. HEYEMANN. *Arch Schiffs Tropen Hyg* 34, 554-61 (1939). -The iodometric method of Kaufmann can be used upon one drop of blood. H. EAGLE

The estimation of iodine in the thyroid glands of albino rats. C. NEWCOMB and G. SANKARAN. *Indian J Med Research* 18, 557-61 (1939). -The gland is dissolved in 10% KOH ashed extd with alc and again ashed. The residue is dissolved in water, and the color which develops upon the addn of H_2SO_4 , Na arsenite, CS_2 and nitroso compared with standards. H. EAGLE

The determination of bile pigments in the blood serum. F. M. MILLER and LUDWIG ENGEL. *Klin Wochschr* 9, 2241 (1939). -The degree of absorption of light of wave length 4203 Å U is an accurate index of the amt. of bilirubin in serum. The diazo reaction used in the van den Bergh test detects only 30-60% the quantity as detd. by the spectrophotometer. The normal value is 0.6-2 mg %. There is no regular correlation between the amounts as detd. by the two types of test. The proportion of bilirubin adsorbed on serum protein when the latter is pptd. may vary from 30 to 70% of the total, moreover contrary to van den Bergh, such adsorption occurs even in sera which give the indirect reaction. In jaundice, there is a parallel increase in the results obtained by the two methods. H. EAGLE

The Meinicke-Klarungs reaction as a rapid micro method for inactivated sera. FRANK MEINICKE. *Klin Wochschr* 9, 2204-5 (1939). -Good results are reported, although the test is not as sensitive as the technic employing active serum. H. E

The use of coal gas in making anatomical preparations in natural color. A. SCHULTZ. *Klin Wochschr* 10, 213-4 (1931). -Tissues may be kept in Jores soln. through which coal gas has been passed, for many weeks without losing their natural blood tinge. Even the color of the organs, such as that of the liver, heart muscle and fat, is preserved by the CO . Jores soln consists of 5% formalin, 2.5% chloral hydrate and 5% of synthetic Carlsbad salts in water. The aeration with CO -contg coal gas must be carried out immediately before the tissue is immersed. After 3-10 days, depending upon the size of the prepn., the tissue is placed in the permanent preserving fluid: 1000 cc. H_2O , 6000 cc glycerol and 300 g Na acetate. For strongly jaundiced organs it is necessary to omit the glycerol and to increase the salt to 500 g per 1000 cc H_2O . H. EAGLE

Determination of cholesterol and lecithin. Evaluation of egg products. J. TILLMANS, H. RIFFART and A. KCHY. *Z Untersuch Lebensmittel* 60, 361-89 (1939). -The methods suggested by other workers are critically discussed. For cholesterol a modification of the method of von Szent Gyorgyi (*C A* 17, 3197) is the most direct and accurate. A soln of not more than 4 mg of cholesterol in 2 cc. of warm acetone is evapd. on a water bath with 1 cc of a 2% soln of digitonin in 80% alc until 1.5 cc remains. After 15 min at room temp. the mixt. is filtered on a sintered glass crucible enclosed in a jacket through which steam can be passed, the ppt. transferred with acetone and twice washed with Et_2O , followed by 3 washes with $CHCl_3$, 2 with Et_2O , 1 with acetone and 1 with cold H_2O (1.5 cc. at a time). The suction is adjusted to filter 1 drop per sec., and the ppt. should always be covered with liquid to protect it from the air. Steam is then passed through the jacket, ten 1.5-cc. portions of H_2O are added, and the washing is finished with the full suction of the pump. The suction and heating are continued while 10 cc. of a clear 1% soln of $K_2Cr_2O_7$ in concd H_2SO_4 is slowly pipetted on the ppt. so as to fill $1/4$ of the crucible. The crucible is washed with 3 1-cc. portions of cold H_2O , sucked well and the total filtrate collected. The oxidation is complete in 1 hr., and 100 cc. of H_2O and 10 cc. of 5% KI soln are added. The liberated I is titrated with 0.1% starch soln. for every 10 cc. of chromic acid taken, allowance being made for the blank titration. The mean of 15 expts. with pure cholesterol showed that the vol. of $Na_2S_2O_3$ soln. equiv. to the chromic acid consumed, divided by 87, gives the cholesterol content in mg. for 0.5-4.0 mg. The factor for phytosterol was 7.9, the m. ps. were 145° and 136°

for the 2 sterols, recrystd from abs EtOH, and 114° and 126° , resp, for their acetates. Egg products are first dried on the water bath with sand, extd in a Soxhlet app with Et₂O, evapd and the residue is dried at 100° and weighed. A weighed portion is then dissolved in warm acetone, filtered, dild to a suitable vol with acetone and the above procedure followed. This gives free cholesterol. Another portion of the ethereal ext. is then sapon on the water bath for 1 hr with a 20% alc soln of KOH, the residue dried mixed with sand and extd as before to give the cholesterol after sapon. Results are given for various egg and pastry products. *Leathin* was detd by various methods, the technic giving the most accurate results consistent with the time consumed is based upon the Jucknack test (*Z. Untersuch Lebensm* 8, 97(1904)). The powd 10 g sample is extd for 3 hrs with hot alc in a Soxhlet app, and the residue left on evapn of the ext is warmed with 15 cc of perhydrol and 5–10 cc of concd H₂SO₄ until brown in color. The cool mixt. is dild to 100 cc, 25 cc neutralized with NH₃ to methyl orange, and the leathin phosphate pptd in the cold in a vol of 60 cc with 5 cc of 1.5% strychnine nitrate soln and 15 cc of a mixt of 33.33 g of NH₄ molybdate in 100 cc of H₂O and 300 cc of dil. HNO₃ contg 200 cc of acid, sp gr 1.4. After 20 min the ppt is filtered off on a weighed crucible, washed with 25 cc of the ice cold mixed reagents (dild 5-fold), and with ice water until free from acid. The wt of the ppt dried at 100° and divided by 39 gives P₂O₅. The method has an accuracy of 0.02 mg for 0.5–1.0 mg of P₂O₅. The contents of H₂O, cholesterol and leathin are tabulated for 21 samples of pastry products contg eggs. A table is given for the calcn of egg content of a sample from the cholesterol content, 58.7 and 71.4 mg per 100 g of ordinary and hard pastry, resp, corresponding with 1 egg per 500 g of flour, and 151.5 and 164.2 mg, resp, for 3 eggs. Storage of samples for a year with the exception of a few home baked pastries showed no loss of leathin or cholesterol. The free leathin content of egg products or pastries is not a reliable means of detg the freshness of the product. C. R. FELLERS

The determination of the concentration of oxidizing agents by means of the residual current. G. TAMMANN and H. THIELE. *Arch ges Physiol (Pfügers)* 226, 694–6(1931).—An attempt was made to devise a method for detg the concn of oxidizing agents in biol materials from the residual current, i , as defined in Nernst's equation: $i = (D \sigma n F / \delta) c$. In this equation D is the diffusion coeff of the depolarizer, n is the electrochem valence of the oxidizing agent, F the electrochem equiv, σ the surface area of the electrode, c the concn of the oxidizing agent and δ the thickness of the diffusion layer at the cathode. With platinumized Pt electrodes roughly approximate results were obtainable for the concn of H₂O₂ in aq solns. The residual current for oxygenated blood was 0.002 amp, as compared to 0.0001 amp for venous blood. A. G.

The determination of the freezing points of small quantities of fluids by means of thermal elements. Experiments with turtle blood. T. MINOSIMA. *Arch ges Physiol (Pfügers)* 226, 794–8(1931).—The f ps of 0.5 to 1 cc quantities of fluids were detd. with an accuracy of $\pm 2\%$ by the use of constantan-iron thermoelements and a sensitive elec arrangement for measuring small p ds. The method is particularly applicable to biol fluids when only small quantities of material are available. ARTHUR GROLLMAN

Influence of formol on the precipitation of serum proteins. MARCEL MASCRE and MAURICE HERBAIN. *Bull soc. chim biol* 12, 978–93(1930), cf *C A* 24, 5052.—Formol affects the pptn of proteins both by forming formol proteins and by increasing the acidity of nonprotein matter present. The latter factor is responsible for the greater effect upon pptn. by alc acetone and salts than by CCl₄COOH. The effect is more marked at lower temps. Pptn with CCl₄COOH includes peptides, and pptn of serum globulins with Na and Mn sulfate includes much of the albumin. C. G. KING

The use of copper and iron salts for the deproteinization of blood. MICHAEL SOMOGYI. *J Biol Chem* 90, 725–9(1931), cf *C A* 24, 3808.—Fe salts can be used instead of colloidal Fe(OH)₃ for the removal of blood proteins in the detn of true sugar. Cu salts are preferable to Fe salts for this purpose and fully the equal of Zn salts in regard to speed and simplicity of technic. For the pptn of plasma or serum proteins Cu is superior to Zn. The technic is described in detail. A. P. LOTHROP

The optimum conditions for the precipitation of casein from human and from cow milk. CHI CHE WANG and AGNES A. WOOD. *Am J. Diseases Children* 40, 787–90(1930).—A clear, edible whey can be prepd from human milk by treating 100 cc of fat-free milk with 33 mg of rennin suspended in 1 cc. of 1 N HCl for 15–30 min at 38° , removing the curd by centrifugation and filtering the supernatant liquid. The max. pptn of casein from human milk by the action of rennin (and HCl) occurs at pH 4.97 and represents 50% of the total N of the whole milk. The max pptn. of cow milk by lactic acid occurs at pH 4.61, representing 80% of the total N. The max pptn by rennin (and HCl) occurs at pH 6.53, representing 73% of the total N. E. R. MAIN

The determination of lead in feces and urine and its significance for the diagnosis of lead poisoning. FRITZ FARTWERTH AND ARTHUR HERTZ. *Arch. Hyg.* 104, 315-25 (1931).—A method is described for the detn. of Pb in feces and urine by which 0.01 mg. of Pb may be detd. The samples are ashed with HNO_3 , the Pb is pptd. first as PbS and finally as PbCrO_4 , and the latter dissolved in HNO_3 and titrated with $\text{Na}_2\text{S}_2\text{O}_4$. The av. values for normal persons who have not worked with Pb are 0.03 mg. Pb per 100 g. of feces and 0.03 mg. per l. of urine; for persons who have worked with Pb but show no symptoms of poisoning, 0.06 mg. per 100 g. of feces and 0.03 mg. per l. of urine; for persons who have worked with Pb and have symptoms of poisoning 0.2 mg. per 100 g. of feces and 0.07 mg. per l. of urine. E. R. MANN

The determination of the p_{H} of phosphate buffer solutions by means of the antimony-antimony oxide electrode. KRISTIN J. FOSBERGER. *J. Lab. Clin. Med.* 16, 411-4 (1931).—The Sb-SbO electrode gives a linear relation between observed e. m. f. and p_{H} values. The p_{H} of phosphate buffer solns. can be detd. with an accuracy of ± 0.01 p_{H} . The electrode is not applicable to the detn. of the p_{H} of blood plasma. E. R. MANN

Chart for computation of daily basal metabolism and percentile basal metabolic rate from spirometer data. CHARLES BARTON. *J. Lab. Clin. Med.* 16, 415-8 (1931). cf. C. A. 24, 322n.—A nomogram-like chart is presented by which the daily basal metabolism and the percentile basal metabolic rate may be computed from the observed values for vol. decrement of O per min., the av. spirometer temp. and atm. pressure. E. R. MANN

The determination of silver in organs and in organic liquids. E. MENEGHETTI. *Riv. Ital. Sci. Sp.* 5, 513-4 (1930).—The organs or liquids are dried and ashed with Na_2CO_3 in a crucible. Complete ashing is obtained by repeated additions of small amts. of KNO_3 . Cooned. HNO_3 is added and evaporated to dryness. For qual. tests, the residue is dissolved with NH_4OH , filtered and the filtrate is tested for Ag (pptn. of Ag_2S with H_2S , AgCl with HNO_3). For the detn. H_2O is added and then the soln. is treated with H_2S for 5 min. and centrifuged; the ppt. is washed carefully 3 times with H_2S water and centrifuged each time. Cooned. HNO_3 is then added to the ppt. and brought to dryness on the water bath. The residue is dissolved in H_2O (up to this point the operations are all made in the original crucible), filtered and the ppt. on the filter washed repeatedly with H_2O . To the filtrate in the dark is then added with caution dil. HCl until complete pptn. of the chloride. The ppt. is allowed to settle and collected on a Gooch crucible, the ppt. being washed first with cold H_2O contg. a small amt. of HNO_3 until the washings are free of chlorides and then twice with alc. It is dried at 100° , then at 130° to const. wt. (all the later operations are carried out as much as possible away from the light). PETER MASTROCI

Electro-ultrafiltration, a method for the determination of the physical state of the inorganic constituents of the serum. RUDOLPH SPRINGER. *Biochem. Z.* 230, 253-5 (1931).—An app. is described for carrying out ultrafiltration with and without electro-dialysis. S. MORGENTHAU

New substrates for use in detecting proteolytic activity. ROBERT L. JONES. *Ind. Eng. Chem., Anal. Ed.* 3, 140-51 (1931).—With a modified method for the prepn. of adsorption systems of egg albumin and dyes a number of reagents can be prepd. which can be used for the detection of proteolytic activity. Two reagents, egg albumin-iodoform and egg albumin-benz. fuchsin, give good results with com. pepsin, trypsin and artificial gastric and pancreatic juices. The prepn. of these reagents is easier than that of previous ones. They can be prepd. in powder form and are stable indistinctly. H. T.

A study of glutathione. HAROLD L. MASON. *J. Biol. Chem.* 90, 479-16 (1931), cf. C. A. 25, 719.—The oxidized form of glutathione prepd. by aeration at p_{H} 7.6, or by treatment with I_2 or $\text{K}_2\text{Fe}(\text{CN})_6$, followed by isolation from alcoholic soln., gives low analyses, as found by Hopkins (C. A. 23, 5477). This is due to the presence in close combination of 2 mols. EtOH , which can only be removed with great difficulty. Heating at $110-20^\circ$ in vacuo leads to loss in wt. but is accompanied by a reduction in the no. of CO_2H and NH_2 groups. The aeration method gives a quant. yield of the oxidized form and is a convenient procedure for its prepn. E. V. THOMAS

Photometric determination of the cholesterol in the serum. E. BRILL AND GUSTAV HÖRNER. *Münch. med. Wochschr.* 77, 142-4 (1930).—Heilmeyer first used the step-photometer, which was designed for the detn. of color in the textile industry and for the detn. of cholesterol. In order to obtain an easier standardization of both halves of the field and to obtain an idea of the light absorption within the various spectral regions, color filters are introduced into the photometer. These are of 2 types (1) 3 filters, red, green and blue, which have a relatively wide penetrable area and (2) a larger no. of color filters which allow very definite, narrowly limited, regions of wave length to

pass. By this app. it is possible to det. the amt. of rays absorbed by a color soln. within the spectral region detd. by the various color filters. From this one can det. the corresponding color value of the soln. Results differed as much as 20% when checked with the Windus method. The colorimetric method gave the same results indicating that other cholesterol like substances give a color reaction. R. C. WILLSON

Determination of Ca, Mg and P in cattle excreta (MORRIS, *et al.*) 7. Device for regulating the quantity and compositions of gas mixtures used for anesthesia (U. S. pat. 1,793,608) 1.

C—BACTERIOLOGY

CHARLES B. MORRIS

The bactericidal power of vosterol. B. E. MONTGOMERY *Proc. Soc. Exptl. Biol. Med.* 28, 491(1930) —Irradiated ergosterol in corn oil neither modified the cultural characteristics nor inhibited the growth of *Bacillus coli* and of *Staphylococcus aureus*. C. V. BAILEY

The present state of chemical study of the tubercle bacillus. E. W. CHARGAFF. *Naturwissenschaften* 19, 202-6(1931) —A review of recent work (cf. Anderson and C., *C. A.* 24, 878, 1131, 25, 128. Sabin, Doan and Forkner, *Am. Rev. Tuberculosis* 21, 200(1930) (cf. *C. A.* 25, 541), Doan, *Proc. Soc. Exptl. Biol. Med.* 26, 672(1929)) with numerous references. B. J. C. VAN DER HORST

Fluoroscopic examination of certain actinomycetes. FRANZ CORTESE *Boll. soc. ital. biol. sper.* 5, 812-4(1930) —The influence of anaerobiosis, pH of culture media and chem. constitution of the media on fluorescence was studied. The fluorescent substance was extd. from the media by weak alc. and purified by shaking with acidified ether. It is an amorphous, odorless, reddish brown material, sol. in alkalis, pyridine, H_2O , MeOH , acetoacetic acid, acidified Et_2O and concd. inorg. acids slightly sol. in weak inorg. acids or acetone, insol. in H_2O or glycerine, benzene and other fat solvents. In 25% NH_3 it shows 4 absorption bands $\lambda = 630-614, 658-676, 660-636, 618-492$. The substance is a pigment belonging to the group of porphyrins. Fluoroscopic examn. showed that *A. albus* and *A. sulphureus* Gasperi are different species of *A. boydii* Harz. PETER MASUCCI

Destruction of sugar in the intestine by the coli group of bacilli. The etiology of pernicious anemia. KOLOMANN V. KNORR. *Folia Hematol.* 43, 32-91(1930) —An organism called *Diplococcus anemiae perniciosae* is described and found present in the mouth and duodenal contents of cases of pernicious anemia. *B. coli* plays a secondary role by splitting carbohydrates and proteins, producing especially compds. with the $\text{NH}_2\text{-OH}$ group. This aids the toxin produced by the diplococcus. JOHN T. MYERS

Toxin production and properties of bacillus Preiss-Nocard. ANGELO ROTTGARDT. *Zentr. Bakt. Parasitenk.*, 1 Abt., 119, 323-33(1931) —The toxin can be pptd. by $(\text{NH}_4)_2\text{S}_2\text{O}_8$ and dried *in vacuo* over H_2SO_4 . Heat destroys it but the dried toxin is more stable than liquid. Cold preserved either liquid or dried toxin. The same is true of the organisms or endotoxin. Air and light have little effect on the exotoxin. Filtration of exotoxin decreases its toxicity about 50%. Toxin concd. at 45° loses a major part of its toxicity but retains its antigenic power. Toluene, ether or mixts. of them have no effect. At a temp. of zero, 5% phenol has no harmful effect. Formalin weakens the toxin, its effect increasing with temp. The pus producing power of the toxin is independent of its toxicity. JOHN T. MYERS

Method for preparing photographs of Petri dish cultures by direct contact printing on photographic paper. J. T. BYRMOLZ AND I. M. LEWIS. *J. Bact.* 19, 195(1930). —The Petri dishes are used in the same manner as negatives and the resulting prints are positives in natural size, the colonies appearing as white objects on a black background. The process is carried out in a dark room, the source of printing light being a 150 w. Mazda lamp at 3 m. from the object, centered directly above the exposing table to avoid cast shadows. The dish, with cover removed, is pressed bottom down upon the sensitized surface of the paper, and exposed. Color and transparency of the agar, its depth in the dish, type of paper, the developer and its temp. influence the time of exposure required. L. G. YFRAGALLO

Resistance of trypanosomes to arseno compounds. M. ADANT. *Compt. rend. soc. biol.* 106, 57-8(1931) —*Tryp. pecaudo* in mice are not rendered resistant to trypanamide by a prolonged or even double contact with trypanamide complex obtained from the liver of mice, which complex, however, according to Levaditi, would be the trypanoxide elaborated by the organism following the injection of drugs such as a toxyl or trypanamide. ALBERT L. RAWLINS

Existence of a specific carbohydrate substance in *B. perfringens*. MARIE E. JIMENEZ. *Compt rend soc biol* 106, 149-1(1931).—The precipitogen of *B. perfringens* in salt water or broth is pptd. by addn of alc contg 1.3% Na acetate; a white, gelatinous sediment is obtained which, on drying in vacuum, yields a yellowish powder almost entirely sol in water. The aq soln. of this powder ppts. strongly in the presence of the serum obtained in the immunization of horses by means of *B. perfringens*; it does not give the biuret reaction, contains a carbohydrate (Molisch reaction pos.), does not reduce Fehling's soln. except after boiling for 1 hr. with 0.5% H_2SO_4 and it contains N.

ALBERT L. RAWLINS

Brucella abortus in Porto Rico. PABLO MORALES-OTERO. *Porto Rico J. Pub Health Trop Med* 6, 1-89(1930).—Sunlight has a deleterious action on cultures of *Brucella abortus*, since cultures exposed to sunlight were sterile in 3 hrs. Ultra violet light from an Alpine sun lamp had a similar effect. The temp was kept low by exposing the cultures on ice, when the temp was allowed to rise the destructive effect of light was even more pronounced. In studying the influence of heat alone on the cultures it was found that a marked diminution of viable organisms occurred in suspensions of *B. abortus* when incubated in the water bath at 42°, but the suspensions were not sterile after 3 hrs. Organisms which survived the exposure to sunlight and ultra-violet light were cultured in broth and injected into mice to test their virulence. The virulence was altered very little if any by the exposure to light. Vaccination with living vaccines as a method for control of Bang's disease was studied, it is believed that this method deserves further investigation. The glucose utilization, the liberation of H_2S and the growth in Huddleson's dye media were studied on 8 strains of *Brucella* and differences were noted. The serological relations of these strains were examd., but with the exception of 2 they could not be differentiated by agglutination or agglutinin absorption. A striking similarity was found between some of the strains isolated from human beings and some of the porcine strains in their action toward CO_2 , their glucose metabolism and the pathol. lesions produced by them in guinea pigs. *B. abortus* seems to be capable of going through a Berkefeld W filter. Sterile filtrates of a broth culture of *Brucella* when injected into mice did not affect them and produced no appreciable pathological lesions in their viscera. Human infection is very low in Porto Rico despite the high cattle infection. Exptl. infection with *B. abortus* in man suggests that the porcine strains are more virulent for man than are the bovine varieties. Only the most virulent bovine strains can infect man through the gastro-intestinal tract. Smaller doses are necessary to produce infection through abraded skin than through the gastro-intestinal canal. Numerous tables, graphs and references are given.

C. SCHWOCK

GARDNER, A. D. *Microbes and Ultramicrobes: Being an Account of the Bacteriophage in its Relations to Bacterial Variation and the Invisible Viruses*. London: Methuen. 2s. 6d., net.

D—BOTANY

THOMAS G. PHILLIPS

Synthetic nutrient solutions for culturing *Ustilago zeae*. EMERY R. RANKEE. *J. Agr. Research* 41, 435-43(1930).—A standard reproducible culture medium is recommended for culturing *Ustilago zeae*. This is a synthetic nutrient soln. which has the following compn., expressed as g./l. of soln.: 0.3 g. K_2SO_4 , 0.1 g. NH_4NO_3 , 0.1 g. $CaCl_2$, 0.1 g. $Mg_3(PO_4)_2 \cdot 4H_2O$, 10 g. dextrose and enough distd. H_2O to make up to 1 l. The reaction of this nutrient soln. before sterilization is pH 7.4 and after sterilization it is pH 5.6. The optimum quantity of dextrose varies from 5 to 10 g./l. depending upon the length of time the culture is inoculated. Maltose is preferable to dextrose for certain physiologic forms of smut. When a solid medium is desired 1.5% of agar can be added.

J. R. ADAMS

Seasonal changes in the catalase activity of conifer leaves. JOSEPH DOYLE AND P. O'CONNOR. *Ann. Botany* 44, 907-15(1930).—In confirmation of the results of Doyle and Clinch (cf. *C. A.* 23, 2459) and contrary to the statements of Burge (*C. A.* 20, 2691, 2700) the authors find that the catalase activity of leaves of *Taxus albertiana*, *T. canadensis*, *Juniperus communis*, *J. rigida*, *Pinus laricio*, *P. strobus*, *P. contorta*, *Abies webbiana*, *J. rigida* and *Picea americana* is in all cases very markedly higher in winter (January) than in May, July or October. Increase in temp. of the plants for short periods during the winter did not alter the catalase activity. In conifers there is no correlation between catalase content and degree of metabolic activity.

J. S. C.

The effect of ionized air on the rate of respiration of fungi. S. R. VAN ASPEREN DE BOER. *Ann. Botany* 44, 999-101(1930).—Expts. with *Phycomyces blakesleeanus* and

Polyporus destructor by the Pfeffer-Pettenkoler method and by a microrespirometer method gave consistently neg. effects upon respiration rate. Ionization of the air was effected by means of Po, the degree of ionization ranging from 200 times to several million times that of normal air. JOSEPH S. CALDWELL.

The resistance to poisons of desiccated plant tissue. WALTER W. ALLEN. *Ann. Botany* 44, 1001-9(1930).—Air-dried material of *Mimulus horum* submerged in abs. alc. for 1 to 13 hrs. was able to grow when thoroughly washed and again placed in the soil. Material placed in abs. alc. 1 hr., then in xylene 1 hr., then washed in alc. and in water, also grew, as did material similarly treated with alc. and acetone or alc. and ether. Submersion in dil. (50%) alc. or in chloroform was fatal. JOSEPH S. CALDWELL.

Further studies on transport in the cotton plant. I. Preliminary observations on the transport of phosphorus, potassium and calcium. F. G. MASON AND E. G. MACKELL. *Ann. Botany* 45, 125-73(1931).—The ringing methods employed in earlier studies of the transport of carbohydrates and N were used (C. A. 24, 5331). In plants deprived of a ring of bark, P, K and Ca accumulate above the ring and diminish below it just as do carbohydrates and N. P, K, Ca and ash constituents ascend the stem mainly by way of the wood, and P and K are returned by way of the phloem to the roots. No evidence of the return of Ca from the leaves was obtained. There is evidence that N, K and P may be transported to the lower portion of the plant in ams. in excess of the needs there, and may there pass into the tracheal sap and again be transported to the leaves. This downward movement of mineral nutrients must affect the rate of entrance of the salts and may in part explain the absence of a marked increase in salt uptake when transpiration is increased. It was possible to reverse the direction of movement of P and of ash constituents by reversing the relative positions of the leafy and leafless portions of the stem. Removal of bolls is followed by marked increases in concn. of P and ash and smaller increases in Ca in the stem and leaves. P, K and some ash constituents travel to the boll along gradients in the phloem, Ca mainly through the xylem. J. S. C.

The effect of chloroform upon the rotation in the internodes of *Nitella*. SUSAN P. NICHOLS. *Bull. Torrey Bot. Club* 57, 153-63(1930).—In continuation of earlier work upon the effect of wounding upon the rotation of protoplasm (C. A. 20, 433) chloroform has been applied to the wall of an internode with a capillary pipet. An area of non-motile protoplasm is produced at the point of application, over which the general protoplasm pushes as rotation continues. The non motile area ultimately resumes movement. Dipping the internode into chloroform for 2-5 seconds, always resulted in cessation of movement after 3-6 hrs. The effect of chloroform is considered to be one of gelation. JOSEPH S. CALDWELL.

Purification and certain properties of the virus of typical tomato mosaic. P. H. BREWER, H. R. KRAVILL, R. W. SWANSON AND M. W. GARDNER. *Phytopathology* 20, 943-50(1930).—A clear, colorless virus suspension was obtained by mincing and pressing affected plants, passing the ext. through the supercentrifuge, discarding the liquid, suspending the residue in dist. water, centrifuging and clearing and decolorizing with slightly acid Al gel. Such a suspension was highly infectious and continued to be so after filtering through a Pasteur-Chamberland F filter or Schleicher and Shull 11/150 collodion filter, but was made non infectious by passage through a Pasteur-Chamberland B filter, collodion filters of 3% or more, or atomometer cylinders. The material retained by the filters was infectious. Of 9 suspensions stored at 36-40° F., 4 became somewhat weakened with age and 1 had lost its virulence after 504 days. The virus was inactivated when alkalinized to pH 7.5-8.5 and reactivated by restoring its acidity. It was not inactivated by acid up to pH 2.46. Heating to 82-84° for 25 min. or to 88° for a shorter time destroys the virus. In cataphoresis tests, the virus-bearing particles accumulate at the pos. pole. JOSEPH S. CALDWELL.

Effect of mineral nutrition on the reaction of wheat varieties to leaf rust. K. D. DOAK. *Phytopathology* 21, 108-9(1931).—Wheat varieties showing various types of reaction to one physiologic form of *Puccinia triticina* were grown in sand cultures with various degrees of excess and deficiency in N, P and K. N increased susceptibility and P and K decreased it. Excess N induced the development of larger primary uredinia, more abundant secondary uredinia and decreased chlorosis. Excess P increased chlorosis and retarded or prevented development of secondary without affecting primary uredinia. Excess K increased chlorosis and decreased the size of primary uredinia. In varieties of intermediate reaction, excess N increased the no. of infection points, while N deficiency, excess P and excess K decreased it. Deficiency of P or of K decreased chlorosis. JOSEPH S. CALDWELL.

Hydrogen sulfide as related to the fungicidal action of sulfur. S. E. Q. MCCALLAN AND FRANK WILCOXON. *Phytopathology* 21, 113(1931).—H₂S is exceedingly toxic to

fungus spores and is evolved when S is applied to spores or to leaves of plants. Germination of conidia and uredo spores of 8 species of fungi were completely inhibited by concns of H_2S ranging from 0.2 to 40 mg. per l., the order of sensitivity being identical with that of these fungi to S. Evolution of H_2S has been demonstrated in 19 fungi and 25 higher plants. Spores of *Sclerotinia* and *Glomerella* produce 5 and 15% of their weight of H_2S in 6 hrs. Indications point to the reduction of S on or within the spores and its initiation by glutathione. The presence of glutathione in spores of *Sclerotinia* has been demonstrated. JOSEPH S. CALDWELL.

The influence of hydrogen-ion concentration and of sodium bicarbonate and related substances on *Penicillium italicum* and *P. digitatum*. RAINYD H. MARLOTH. *Phytopathology* 21, 169-69(1931).—Citrus fruits are dipped in a 35% soln. of $NaHCO_3$ prior to packing to prevent decay by blue and green molds (*Penicillium italicum* and *P. digitatum*). An attempt was made to det. whether the action is specific or has a general physiol. basis. The fungi were cultured on modified Duggar's soln. to which an autoclaved orange ext. was added in an app. which permitted daily renewal of the medium. *P. italicum* grew nearly equally well over a pH range of 2.9 to 6.5, *P. digitatum* over a range of pH 3.0 to 6.0. Germination of spores in hanging drop cultures was inhibited at pH 1.5-2.5 and excellent at 3.0 to 5.3, falling off at pH 7.0-9.0. The Na ion markedly inhibits germination at concns of 15,000-20,000 p.p.m.; the K ion does not. Na tetraborate is more toxic to *P. digitatum* and $NaHCO_3$ is more toxic to *P. italicum*. The effect of the wash treatment in preventing decay is due to the fact that the film of salt left upon the fruit forms a sat'd solution of Na_2CO_3 about the developing germ tufts which is toxic thereto. JOSEPH S. CALDWELL.

Enzyme formation in *Penicillium glaucum*. I. SAKIMABURO WADA. *Arch. Sci. Med. Univ. Imp. Kioto* 13, 124-41(1929).—A study of the relationship between the concn. of the nutrient medium and the enzyme content of *Penicillium glaucum*. When this mold is grown on a protein free, carbohydrate free medium (salts and glycerol), it contains desamidases, urease, a glucosamine-splitting enzyme, nuclease, pepsin and trypsin like enzymes and crepsin. It does not show the relation between the nutrient medium and enzyme content exhibited by bacteria. II. *Ibid.* 145-57.—In the mold cultivated with the above medium lactase and a glycolytic enzyme were absent but these are also lacking in a carbohydrate-contg. medium. In the culture liquid, a nuclease, pepsin, glycogenase, invertase, a glucosidase, a lipase and a hippuric acid-splitting enzyme were present. III. *Ibid.* 158-63.—A comparison of the enzyme content of the mold when grown on a medium contg. protein and carbohydrate (glucose, meat ext. and peptone) with that above shows that the pepsin, trypsin and lipase content is much decreased when these foodstuffs are absent from the medium. II, J. D. J.

Chlorophyll formation. KURT NOACK AND WILHELM KIESLING. *Z. anorg. Chem.* 44, 67-6(1931). cf. C. A. 25, 721.—Protochlorophyll, the precursor of chlorophyll, is present in small amt. in seedlings germinated in the dark but is more easily obtained from the inner membrane of the seed coat of pumpkin seeds. It yields a series of derivs. analogous to those prep'd by Willstätter from chlorophyll. Like chlorophyll it contains 3 carboxyls, one of which is esterified with phytol, another with Me, while the 3rd occurs in lactone or lactam linkage. Removal of Mg by acid treatment yields protopheophytin contg. the phytol, Me and lactam groupings of the parent substance. Esterification of this with 30% HCl in MeOH replaces the phytol by Me and esterifies the lactam carboxyl, yielding protophytochlorin trimethyl ester, m. 234-5°. The opening of the lactam ring causes a marked change in the spectral behavior of the protopheophytin as shown in the intensity sequence of the absorption bands. Partial sapon. of protopheophytin should yield protopheophorbide, but this method is unsatisfactory because the concn. of alkali required leads to anhydridization. Protopheophorbide was obtained, however, by partial sapon. of the above trimethyl ester. The substance contains 1 CO_2H , 1 CO_2Me and the lactam group, and is analogous to pheophorbide. The dimethyl ester, methylprotopheophorbide, in which the lactam group remains intact, is best prep'd. by reduction of methylpheophorbide a from chlorophyll by means of Fe and 80% CH_3OH . Its spectrum is midway between that of protopheophytin and that of phyloerythrin. Both protopheophorbide and its Me deriv. yield the same trimethyl ester identical with that obtained from protopheophytin. Derivs. of the protochlorophyll series differ from the corresponding chlorophyll derivs. in that they contain 1 less O. This O, which is a structural part of the chlorophyll, is present in a linkage as yet undet'd. Phyloerythrin, the chlorophyll deriv. present in beef bile, is closely related to protopheophytin. Like the latter it yields a green pigment when Mg is introduced by the Grignard reaction with $p-Me_2NC_6H_4MgI$. The pigment thus obtained resembles protochlorophyll spectroscopically in the same manner that phyloerythrin resembles proto-

pheophytin, ϵ , in a displacement of the absorption bands toward the violet. Derivatives of the phyloerythrin series are isomeric with the corresponding protopheophytin derivatives. The phyloerythrin used in this work was prepared from 900 l of beef bile. It is neither an ester nor a free acid but an intramolecular anhydride, and may be converted into an acid by treatment in the cold with KOH in $\text{Me}_2\text{C}(\text{OH})_2$. The accompanying spectral change is similar to that observed in the conversion of protopheophytin to the protophytochlorin trimethyl ester. The product is probably a tricarboxylic acid in which 1 CO_2H is anhydridized. With CH_3N_3 it yields a trimethyl ester, m 233-3°, isomeric with the trimethyl ester of protophytochlorin. Phyloerythrin, like protochlorophyll, is a reduction product of chlorophyll. It gives a secondary series of derivatives in which a splitting off of CO_2 has occurred. Esterification of phyloerythrin with MeOH-HCl yields a mono- or a mixt. of mono- and diesters according to the intensity of action. Sapon of the ester gives a tricarboxylic acid contg. 1 Me, and thus when esterified yields a trimethyl ester, m 205°. Phyloerythrin thus contains an acid anhydride and a lactam group, the latter probably identical with that of chlorophyll. Chlorophyll formation represents a photo-oxidation of protochlorophyll. The addition of O yields chlorophyll a, which in turn is the precursor of the more highly oxidized chlorophyll b. Phyloerythrin formation is the reverse process and consists not only in reduction, but also in Mg cleavage, sapon and anhydridization. Probably phyloerythrin is a biol. transitional stage between leaf and blood pigment. A. W. DOX

The copper and manganese content of some vegetable tissues. MARIA VAN LEEUWEN. *Ann. physiol. physicochim. biol.* 6, 178-81 (1930).—Cu and Mn are present in approx. the same concn. (0.2-2.5 mg % averaging around 1) in the leaves of rhododendron, nasturtium, ivy, vine, hops, poplar, fir, etc. H. LAGLE

Saponin from spinach. O. DAFERT. *Z. Untersuch. Lebensm.* 60, 409-0 (1930).—Very little saponin occurs in the stem or leaf, and none at all in the seed of spinach. The hemolytic index of the root, because of its saponin content, is 1,000 for sun-dried roots, and 1,400 for roots dried in an oven at 100°. The index rises from 1.85 for a plant 3 weeks old to a max. of 1,800 after 7 weeks, and then falls to 1,280 after 20 weeks. The saponin is extd. from the scraped and sun-dried root in a 5-fold vol. of 70% EtOH for 2 hrs., followed by extn. for 1 hr. with a 2-fold vol. of EtOH. The exts. are filtered, the residue is pressed and the crude saponin then separates from the filtrate after 1 day in the ice-chest. It is filtered off, washed with abs. alc., the filtrate evapd. to 1/2 its vol., cooled and an addnl. yield of saponin is thus obtained. The total quantity is dried, extd. with LiOAc , and the saponin recrystd. from hot 90% EtOH as pointed white needles with a hemolytic index of 1,100,000 and m 260-2°. The properties of saponin are given. Hydrolysis with H_2SO_4 in 50% LiOH, and treatment with charcoal gave a white amorphous sapogenin, $\text{C}_{27}\text{H}_{48}\text{O}_5$, m 208-12°, sol. in alc. and EtOAc and insol. in H_2O , CCl_4 or Li_2O . C. R. FELLERS

Vegetable rennet. D. MILNE. *Rept. Operations Dept. Agr., Punjab, Year ending June 30, 1929*, Pt. I, 49-0 (1930).—The ripe berries of *Hibiscus coagulans*, which grows wild on the Punjab plains, yields an ext. which is rich in rennet. The crude ext. imparts a brown coloration and a bitter taste to milk and is unsuitable for use in the prepn. of cheese. By repeated pptn. the active principle is obtained free from these objections and in a very pure condition, one part of the prepn. being sufficient to coagulate 1000 parts of milk in 1 min. K. D. JACON

Decomposition of salts of organic acids by the mold fungus *Aspergillus fumigatus*. W. THIES. *Ber.* 64B, 214-8 (1931).—*Aspergillus fumigatus* decomposes sugar, forming not only H_2CO_2 and $\text{H}_2\text{C}_2\text{O}_4$ as end products, but also citric (I), fumaric (II) and gluconic (III) acids. If other acids are formed as intermediate products, the fungus should decompose these acids in sugar-free cultures to I or II. Two series of expts. were carried out; one with the Ca salts of I, II and III, and of saccharic, succinic, tartaric, malic and lactic acids, the other with the K, Na and NH_4 salts of the same acids. I and II were included in order to det. whether they are exclusively end products, or whether they are further utilizable by the fungus. In the first series H_2CO_2 was the end product of the decompn., as shown by the formation of considerable quantities of CaCO_3 ; in the second, the end product was $\text{H}_2\text{C}_2\text{O}_4$. The results indicate that the nature of the cation affects the change: if the cation is bivalent, the end product is H_2CO_2 ; if univalent, the end product is $\text{H}_2\text{C}_2\text{O}_4$. The expts. show clearly that there is no support for the theory that in the decompn. of sugar to I or II, other acids are formed as intermediate products. LOUISE KELLEY

The influence of environmental factors on acid formation by *Aspergillus fumigatus*. W. THIES. *Zentr. Bakt. Parasitenk.*, 2 Abt., 82, 321-47 (1930); cf. preceding abstract.—The influence of various factors as, neutralization with Pb or BaCO_3 , O tension, character of the medium and pH , had little influence on acid production. Gluconic and citric

acids are always formed, oxalic and fumaric acids sometimes result. Other acids are not formed. JOHN T. MYERS

Presence of rutoside in the leafy stems of *Bupleurum falcatum* L. J. RABATE. *Bull soc chim biol* 12, 974-7(1930)—Extn with alc gave 4.3 g per kg dry wt.

C. G. KING

Wall structure and mineralization in coralline algae. L. G. M. BAAS-DECKING AND E. WAYNE GALLIER. *J Phys Chem* 35, 467-79(1931)—The optical properties of the cell walls of the algae, *Corallina officinalis* Lamark and *Amphiroa dorbigniana* Decaisne were studied in the hopes of detg their chem and phys nature. The data show that the cell walls of these 2 organisms are made of a non-cellulosic material, probably a pectin like substance, and are birefringent in longitudinal and isotropic in cross section. The fibers are built up of tangentially arranged, elongate lamellae interspersed with concentrically arranged interstices, both of which are small in relation to the wave length of light. Calcite is the only mineral deposited by the living organism, the individual crystals of which are only a few tenths of a micron long, and arranged with the c axis perpendicular to the longitudinal axis of the fiber. The deposition of Mg is a secondary phenomenon.

W. LEAHY

Investigations on the nutrition of fruit trees. Some effects of deficiencies of nitrogen, potassium, calcium and magnesium, with special reference to the behavior of certain varieties of apple trees. M. B. DAVIS. *J. Pomology and Hort Sci* 8, 316-44 (1930)—Bramley seedling, Worcester Permain and Allington Pippin apples were grown under controlled conditions in pots in sand and in soil and given differential nutritional treatments. Omission of N resulted in restricted shoot growth, yellowed leaves and delayed bud break in spring. Lack of K_2O caused early partial defoliation and later retention of the remaining foliage, leaf scorch and much earlier bud break. P_2O_5 omission produced earlier defoliation, very restricted shoot growth, typical bronzing of the foliage and delayed bud break. Absence of CaO resulted in increased shoot growth and larger leaves, which later broke down either in the center or along the margins. Omission of MgO was characterized by reduced shoot growth in most cases, earlier defoliation and marked breakdown of foliage as brown spots in the center and margins of the leaves. The omission of an element always resulted in a lower percentage of that element in the ash and dry matter of the young shoots and leaves. Reduction of K_2O was associated with increases of CaO and MgO , low CaO was accompanied by high K_2O , high MgO and low P_2O_5 , low MgO was associated with high CaO , low P_2O_5 and low K_2O . A high degree of correlation existed between symptoms exhibited and amt of the related element in the ash and dry matter.

A. L. MEIERING

Prolonging the longevity of rice seed. G. SAMPIETRO. *Giorn. viticoltura* 21, 1-5 (1931)—Rice seed, dried to a moisture content of 5% and kept in an atm of N, gave a 99% germination after 8 years. Seed kept in CO_2 , in vacuum or in air did not germinate at all when dried to a normal 13% or to a 5% moisture content; likewise seed kept in N when contg 13% moisture.

ALBERT R. MERZ

The influence of iodine on the reproduction of yeasts. J. KOOLJANS. *Zentr. Bakt. Parasitenk.*, 2 Abt., 82, 347-53(1930)—Minute amts. of I (about one ten millionth part) have a slight growth stimulating effect on some strains of yeast. There is no relationship between I and bios.

JOHN T. MYERS

The pigment of the watermelon. L. ZECHMEISTER AND P. TILSON. *Ber* 63B, 2881-3(1930)—The pigment consists of a mixt. of lycopin and carotene (8-10 parts of the former to 1 of the latter in melons from southern Hungary). In small scale expts 0.2 g of the mixt was obtained from fourteen kilograms of the fresh melons, while 150 kg of the melons yielded 1.3 g lycopin and 0.07 g carotene, and another 0.07 g of the latter remained in the mother liquors.

C. A. R.

Permeability of I on some economic plants (MALHOTRA) 15. Sabineroin of *Salix cinerea* L., its identity with piceoside (RABATE) 10.

E—NUTRITION

PHILIP B. HAWK

Animal experiments on the significance of increased vitamin administration in the therapy of tuberculosis. W. FRANSTIEL AND B. SCHARLAU. *Beitr. klin. Tuberk.* 73, 351-72(1930)—The authors used not only primary infected animals but also superinfected immune animals which had been given the preliminary infection with low virulent human tubercle bacilli. In rabbits vitamin B concentrate contg in 1 cc the equiv. of 100 g of yeast, and vitamin D (ergosterol, viganol) revealed no therapeutic effect. Vitamin A (cod liver oil), however, had a favorable influence on some of the animals.

Therapeutic action was noted with a coincident peroral administration of yeast and raved egg sterol. With weak infections in immune animals there was a complete lack of tuberculosis following the use of vitamins E and D, while with larger infection there occurred a favorable type of disease. A similar, though smaller, effect was noted in primary infected animals. A decidedly less effect resulted when vitamin D only was not given perorally but by inhalation. Primary infected animals were unfavorably affected by overdoses of vitamins E, D and A. In children wt control can prevent overdoses of vitamin D, and the danger of a B hypervitaminosis need not be feared. H. J. C.

Vitamin A and carotene. NORMAN S. CARVER. *Nature* 126, 603, (1930).—The administration of 1 mg. doses of carotene daily to chickens on a vitamin A free diet cured the vitamin A deficiency and enabled normal growth to take place. The livers of these birds gave strong positive tests for vitamin A. It is concluded that carotene may serve as a precursor of this vitamin. H. J. PETERS, JR.

The significance of the ultra-violet range for the protective substance against rickets. WALTER HILSON and LOUIS KELLNER. *Can. Med. Ass. J.* 10, 171, (1931).—Ergosterol, has an absorption band at 3.2 μ (ultra red) following irradiation with ultra violet light the ergosterol develops a new weak absorption band at 0.9 μ , while the normal band increases in intensity. Ergosterol irradiated with ultra violet light, has 2 absorption ranges in the ultra violet range at 290.80 and 293.90 m μ . Both are increased by subsequent ultra red irradiation. In the case of the former, the increase signifies a decomposition of the protective substance. In the latter it signifies the formation of a new substance. The autoxidative factor of ergosterol irradiated with ultra-violet light is destroyed by the ultra red. H. PETERS.

Importance of sugar as the fuel of life. F. FLETCHER. *Arch. Pharm.* 269, 9-22 (1931).—An address. W. O. R.

Detection of vitamin A in irradiated food products. H. MARTIN and R. NORTON. *Arch. Pharm.* 269, 22-4, (1931). The results obtained with 3 series of rats in an experiment of the products Novotiron and Fertigan are graphically shown in connection with the keratomalacia laboratory and later diagnosed by feeding with vitamin A free products and the above-named irradiated foods. W. O. R.

Influence of the canning industry in the changing dietary. F. F. KOWAL. *J. Am. Dietetic Assoc.* 6, 123, 9, (1930).—Canned foods have 3 distinct advantages from the standpoint of their vitamin: (1) they are always canned promptly after harvesting; (2) oxidation, which destroys the vitamins, is not permitted to take place in canning as it does in open kettle cooking; and (3) after canning, the vitamins in canned foods are apparently stable and may be stored for long periods with insignificant or no loss. For example, canned tomatoes are richer in vitamin A and B than orange juice and are only slightly lower in vitamin C. Canned fish products of all descriptions contain an unusual amt. of I. The diet, even though complete in its chemical constituents, will not produce its optimum effect unless eaten with relish and enjoyment. C. R. FLETCHER.

Influence of fertilizer treatment on the vitamin A content of spinach. HANNAH H. HOSKINS and R. ANTON DUTCHER. *Penna. Agr. Expt. Sta., Bull.* 268, 9, (1930); *3d Ann. Report* (Abstract).—Fourteen samples of Rhode Island spinach, dried and stored under CO₂, were assayed for vitamin A content. The spinach was grown under various fertilizer treatments. No significant results were obtained except that where spinach is grown with Mn as a limiting soil nutrient, less vitamin A is found. This indicates that pigmentation, metabolism and vitamin A synthesis are correlated. The work is being continued. C. R. FLETCHER.

"Luxury consumption" and feeding. R. FOWLER and T. H. TUNT. *Arch. intern. physiol.* 31, 363-37, (1929). *Physiol. Abstracts* 15, 245.—In overfeeding, rabbits and pigeons do not make a "luxury consumption," and in underfeeding the breakdown processes of the body are cut down correspondingly. C. G.

Physiological ketogenesis in herbivores. R. H. FLETCHER. *Arch. intern. physiol.* 31, 363-412, (1929); *Physiol. Abstracts* 15, 245, (1930).—In herbivores there is an abundant excretion of keto bodies on any diet or in transition, these substances are of both endogenous and exogenous origin. Proteins are more ketogenic than carbohydrates, which are in turn more so than fats. The addition of HCl to the food diminishes the excretion of "keto bodies," but there is no relation between the excretion and the total acidity or alkali of the urine. Variations occur in any one animal on a given diet, and also in different animals. C. G.

Mineral feeding experiments. Iodine, lime and salt-lick questions. R. W. SIMPSON. *New Zealand J. Agr.* 42, 1-23, (1931).—The I content of feeds and hay was increased about 10-fold by the application of small amounts of I salts to the soil. The

percentage increase in the growth of rabbits receiving I either in the form of inorg. salts or in the form of foods high in I was distinctly greater than that of the control animals. Addn. of lime, alone or in the form of salt lick, to the diet of rabbits did not cause enlargement of the thyroid gland. K. D. JACOB

The vitamin D content of grasses varying in origin and fertilization. ARTHUR SCHEUNERT and JOHANNES RESCHKE *Tierernähr* 2, 262-9 (1930) —A no. of green grasses of varying origins were tested for their vitamin D content. Different types of grass have different vitamin D contents. The same type of grass may show a widely varying vitamin D content. The reasons for this cannot be explained. No decisive influence seems to be attributable to the use of artificial fertilizers. The conditions of the soil presumably play the most important part. J. C. JURKENS

A new cereal mixture containing vitamins and mineral elements. FREDERICK F. TISDALL, T. G. H. DRAKE and ALAN BROWN *Am J Diseases Children* 40, 791-9 (1930) —The prepn. of a cereal mixt is described, composed of wheat meal, oat meal, corn meal, wheat germ, bone meal, dried brewer's yeast and alfalfa. The mixt. supplies the minerals and vitamins which are lacking in the ordinary milled, part grain cereal. The Fe content of the mixt. is 0.024%, the Cu content, 0.0013%, and the Ca content, 0.78%. Vitamin A is supplied by the wheat germ and the alfalfa, vitamin B₁ and B₂ by wheat germ and yeast and vitamin I by the wheat germ. Vitamin D may be secured by irradiation either of the wheat germ or the yeast. The mixt. is palatable and does not irritate the intestinal tract. E. R. MAIN

Milk sugar in infant feeding. A study of the effects of the routine use of milk sugar in infant feeding. B. WINSTON JAAVIS *Am J. Diseases Children* 40, 993-9 (1930) —Lactose given to infants in 10 g. quantities at 3-day intervals is well tolerated when methods are used to give the other elements of the diet their max. digestibility. Infants fed with lactose resemble breast fed infants in firmness of tissue and resistance to infection. When compared according to age-wt., they possess more living tissue than infants fed on vegetable sugar, and are less susceptible to sudden dehydration with acidosis as a result of parenteral infection. The use of lactose is desirable as a source of galactose for the development of the central nervous system. E. R. MAIN

The carbon dioxide absorption curve of infants fed lactic acid, hydrochloric acid and boiled cow milk. DANIEL C. DARROW *Am J. Diseases Children* 40, 1016-23 (1930) —The CO₂ absorption curves for children fed either boiled whole milk or milk to which lactic acid, or both lactic acid and HCl have been added are essentially normal. The same changes in diet produce no significant changes in the bicarbonate concn., lactic acid concn. or p_H of the blood. The p_H of the venous blood, in infancy, is 7.37 ± 0.05. The venous CO₂ tension is 41 ± 5.0 mm. of Hg. The hemoglobin exerts the same influence on the slope of the CO₂ absorption curve as in adults. E. R. MAIN

Recent progress in poultry nutrition. GREGORY CHOMKOVIC and JAN PODHRADSKA *Russ Arch. Landw., Abt. B., Tierernähr. Tierzucht* 2, 27-214 (1930) —A review. Seven pages of references are appended. W. GORDON ROSE

Banana—a challenge to chemical investigation (VON LOESSCKE) 12. Vitamins and palm oil in margarine (EPSTEIN) 27. Determination of Ca, Mg and P in cattle excreta (MORRIS, et al) 7.

BERG, RAGNAR. Eiweissbedarf und Mineralstoffwechsel bei einfachster Ernährung. Leipzig: Huzel, 239 pp. M 10.50, linen, M 12.

F—PHYSIOLOGY

E. K. MARSHALL, JR.

Citric acid content of animal fluids (cerebrospinal, follicular and amniotic fluids and aqueous humor). I. I. NITZESCU and I. D. GEORGESCU *Compt rend* 190, 1325-7 (1930) —The citric acid content of the following animal fluids has been detd. by Thunberg's methylene blue test: human cerebrospinal fluid 0.075-0.080 g. per l.; of dog, 0.090-0.100 g. per l.; aq. humor of oxen, 0.012 g. per l.; follicular fluid of cows, 0.040 g. per l.; human amniotic fluid, 0.025-0.027 g. per l. B. C. A.

Occurrence of metallic elements in biological material. W. B. S. BISHOP *J. Cancer Research Comm. Univ. of Sydney* 1, 212-53 (1930) —The elements of groups I and II in the periodic table are widely distributed. Group III is the poorest as regards biol. occurrence, the presence of Al in animal tissues being much in dispute. Pb (group IV) is the 1st toxic element occurring in tissues. As and V are biologically the most interesting members of group V, the latter being but slowly eliminated after injection. The carcinogenetic action of tar does not depend on its As content. Mn is found in small

amt. in animal tissues but is not related to sarcomata or carcinomata. Cu is connected with hemoglobin formation, while Zn may play some part in reproduction. B. C. A.

Fibrillar structures in the albuminous layer of the egg of the fowl. G. C. HERINGA AND S. H. VAN KEMPE VALK. *Proc. Acad. Sci. Amsterdam* 33, 530-2 (1930).—The albuminous layer surrounding the egg yolk consists of a series of strata, the outer layers being clear and transparent, while the rest is opaque. The membranes are of fibrillar structure. The fibrils are dissolved on digestion with pepsin and HCl at 37°, when the albumin of a hard boiled egg is dissolved completely. In a raw egg the membranes are dissolved less readily than is the remainder of the albumin, and when sections are prepared from a frozen egg the membranous structures remain in connection even after digestion. The membranes are completely sol in alk. trypsin. Evidence is obtained that the fibrils are composed of a keratine substance. B. C. A.

The absorption of ultra-violet light by the liquids of the organism. J. CLUZET AND T. KOFMAN. *Compt. rend. soc. biol.* 103, 1123-5 (1930).—Serum, plasma, urine, bile and aq. humor (ox) are completely opaque in a layer 15 mm thick to ultra-violet light (3000-2300 Å). Cerebrospinal fluid, however, has an opacity relative to H₂O of 68.6. The opacities of increasing dilns of these liquids show important differences. The electrolyte contents of bile and urine are responsible for as much absorption of ultra-violet light as the proteins of serum or plasma. The opacity of defibrinated blood at dilns. of 1/100-1/1000 is much greater than that of plasma for the same dilns.

B. C. BAUMSTETTER.

Acetylcholine in beef blood. Comment on the paper of J. Kapfhammer and C. Bisthoff, FRITZ WREDE AND WERNER KEIL. *Z. physiol. Chem.* 194, 229-31 (1931).—Kapfhammer and Bisthoff (*C. A.* 25, 127) have reported the occurrence of acetylcholine (28.8 mg per l) in beef blood, but neglected to verify their findings by physiol. tests on isolated intestine. Exts. prep'd by the method of K. and B. failed to show any physiol. action on guinea pig intestine in Tyrode soln. An ext. which should contain 2.9 mg acetylcholine per cc according to K. and B. gave no appreciable contraction with quantities of less than 5 cc. Acetylcholine added to the blood in quantities up to 30 mg per l disappeared completely during the process of purification. On the other hand, treatment of acetylcholine-HCl salt with 10% CCl₃CO₂H during 12 hrs. did not impair its physiol. action. It appears that the acetylcholine, originally present or subsequently added, is carried down in the protein pptn., and does not appear in the ext. as claimed by K. and B.

A. W. DOX.

Choline in the uterus and its relation to labor pains. ERICH STRACK AND ADALBERT LOESCHKE. *Z. physiol. Chem.* 194, 269-76 (1931), cf. *C. A.* 23, 5493.—Although the human placenta contains 180 mg of choline per kg, no choline could be found in the gravid uterus. The av. choline content of the non gravid human uterus was 47 mg per kg. Beef and pig uteri, on the other hand, showed practically the same choline content in the gravid as in the non gravid state, approx 160 mg. No physiol. relationship could be demonstrated between choline content of the uterus and the onset or severity of labor pains.

A. W. DOX.

Regulation of the secretion of insulin, with observations on the regeneration of the islands of Langerhans in the pancreas. G. JORNS. *Klin. Wochschr.* 8, 2319-22 (1929).—An inconclusive, highly theoretical discussion.

H. EAGLE.

The hormones of the anterior lobe of the hypophysis. L. BERNHARD ZONDER. *Klin. Wochschr.* 9, 245-8 (1930); cf. *C. A.* 23, 4733.—The hormone of the anterior pituitary lobe is the true sex hormone. An ext. (prolan) injected into the infantile rat causes (1) ripening of follicles, ovulation and rut, (2) the appearance of blood points in the ovary, and (3) formation of corpora lutea. Prolan is not the same as the growth-promoting ext. of Evans, and although it contains prehypyon, the substance which regulates the sp. dynamic action of foods (Kestner), its hormonal effects upon the sex glands are not due to this substance. The sex hormone in prolan contains at least 2 distinct substances one which initiates follicular ripening, uterine hyperemia and mucosal changes in the vagina of infantile white rats 4-5 weeks old (prolan A), and another which causes the formation of corpora lutea in 3-4 weeks old female mice, and which has no effect upon the uterine or vaginal mucosa (prolan B). Folliculin is formed both by the follicles and by corpora lutea in women, but only by the former in animals. In the animal prolan A would mobilize folliculin, which induces rut, prolan B would induce luteinization. The formed corpora lutea would then induce the hypersecretion of the uterine and vaginal mucosa by their hormone, which Corner calls lutin. II. Follicular ripening hormone (prolan A). Menopause. Castration. *Ibid.* 393-6.—Prolan A can usually not be demonstrated in the urine of non-pregnant women unless it is concentrated more than 5 times before extn. At the onset of clauacterium there is an enormous

increase in the excretion of folliculin, up to as high as 200 units per 1000 cc. of morning urine, to be contrasted with the 10 and 30 units normally excreted during the intermenstrual and premenstrual period, resp. There is at this time a polyfolliculin amenorrhea. The second stage begins weeks or months later and is characterized by a complete absence of folliculin from the urine with accompanying vasomotor disturbances. Finally, there is increased excretion of prolan A which can be demonstrated in the urine of 25% of women at the menopause. Following removal of both ovaries, prolan A promptly appears in the urine in 75% of the cases. Even after Röntgen ray castration, it appeared in 4 out of 10 cases. Z. suggests that it is excreted in the urine only when the ovaries, its elective site of action, are removed. Its increased formation following castration may be due to the removal of an ovarian inhibiting substance. It is significant that the hypophysis undergoes histological changes at the menopause or following castration.

H. FAGLE

Energy metabolism and its clinical significance (formation of lactic acid and oxygen consumption, function and metabolism, diet.) S. G. ZONDER AND F. MATYAS. *Klin Wochschr* 10, 11-7(1931).—The recent evidence tending to show that the formation of lactic acid is not the actual cause of muscle contraction is summarized. (1) Acetylcholine and other substances cause contraction without the release of lactic acid. (2) In tonic contractions, the formation of lactic acid follows the beginning of contraction. (3) Lactic acid production can be entirely inhibited by bromoacetic acid, e. g., the possible role of phosphocreatine is discussed. In muscles under tension there are structural and functional changes which may explain the anomalous increase in lactic acid. Similar aerobic glucolysis is observed in other conditions involving nonspecific disturbances in metabolism (Kasedow's disease) as also in the metabolism of old leucocytes and mature erythrocytes. Deviations from the usual course of metabolism, in general, may therefore be due to non specific cellular injury. Thus, lactic acid is not the cause, but a product, of muscle fatigue. The authors discuss the relation of these problems of muscle metabolism to clinical conditions.

H. FAGLE

The relationship of fat in its passage through the placenta (contribution to the physiology of the placenta). WERNER DICKENBACH AND HANS RUPP. *Klin Wochschr* 10, 63-4(1931).—The fetus receives its fat directly from the maternal organism, unchanged by its placental passage. Thus, if pregnant rabbits are fed with linseed oil, containing an oil with an I no. of 115-175 (normal rabbit fat = I no. of 80-90), the fat of the fetus is found to have exactly the same I no. as that of the mother. There is no real evidence for a splitting of fat in its passage through the placenta, such as occurs in the intestinal wall. Indeed, if the methyl esters of oleic acid, and the amyl ester of palmitic acid are injected into a pregnant rabbit intravenously, they can be found as such in the fetal tissue, but no free amylalcol or McOil is demonstrable, indicating that there had not been any splitting of the fat in its placental passage.

H. FAGLE

Pyruvic acid in the metabolism of animal cells. BRUNO MANDAL, MILLY BAUCH AND FRIDA SKILITZ. *Klin Wochschr* 10, 118-9(1931).—Pyruvic acid increases anaerobic fermentation by body cells more than 100%, causing them to approximate glycolytic activity of cancer cells. A simple test for pyruvic acid is described, which is based on the fact that it prevents glyceraldehyde from inhibiting the fermentation by Jensen sarcoma, a test which detects several thousandths of a mg. Using this test, as well as that of Warburg, the authors find that cancer cells under normal conditions form very large quantities of pyruvic acid. It is also present in mammalian sera 10 mg. % in the serum of white mice, and somewhat less in rabbit, rat, guinea pig and human serum.

H. FAGLE

The solubility, dissociation and tension of carbonic acid in urine. LAITZ MAIVRE. *Klin Wochschr* 10, 119(1931).—The soly. coeff. α (34°) varies between 0.44 and 0.51, the apparent dissociation const. P (38°) is 5.8-6.3. The tension is very variable, ranging between 16.7 and 212 mm. Hg in 20-odd urines, obviously, however, it is much greater than it can possibly be in the tissues.

H. FAGLE

The occurrence of very small quantities of mercury in the urine and stools. Its diagnostic significance and causes. PAUL BOHNECK. *Klin Wochschr* 10, 149-53(1931).—34 of 75 normal people excreted as high as 10 μ g Hg daily, despite the absence of any known exposure to Hg. People with recent dental amalgam fillings excrete considerably more. Almost all foods contain Hg, the total quantity in the daily intake being about 5 μ g, corresponding to the urine output.

H. FAGLE

Activating split products and hormones of the tissues. WOLFGANG WEICHARDT. *Klin Wochschr* 10, 163-4(1931).—Tissue extracts injected into mice increase the performance of its muscles, in larger doses they cause paralysis. Non specific toxic processes may cause the liberation of these tissue hormones in the body, as shown by expts. upon

hearts. Very small quantities have no effect upon normal times, but do increase the performance of fatigued organs. Thus, 1 mg. of a hydrolysis product of the skin has a definite effect upon the contraction of a fatigued heart. H. KAUH.

The specific gravity of blood and its constituent parts. C. ORSTADEN. *Klin. Wochenschr.* 10, 100-2(1931).—The sp. gr. of whole blood is 1.05-1.06, that of plasma 1.025-1.030. In *perniciosa anemia*, the sp. gr. of the erythrocytes proper is increased from a normal value of 1.085-1.091 to 1.095-1.1. H. KAUH.

The time relationship of the formation of lactic acid in muscle contraction. OTTO MEYERHOFF. *Klin. Wochenschr.* 10, 211-3(1931). M. confirms the finding of Lehmann that the supposed const. $K = \text{tension (g.)} \times \text{duration of tetanus (seconds)} \times (\text{cm. of muscle length}) / \text{lactic acid formed (g.)}$ is larger if the muscle is killed immediately after the tetanic contraction than if 2-3 min. are allowed to elapse, i. e., that there is a post-tetanic formation of lactic acid. However, he is unable to find that the const. $K = \text{tension (g.)} \times \text{muscle length (cm.)} / \text{lactic acid (g.)}$ for single contractions is dependent upon the time interval between stimuli, over a wide range (1-10 sec.). This late formation of lactic acid may play an important role in making possible the anaerobic resynthesis of creatinephosphoric acid which is an endothermic reaction. H. KAUH.

Growth and development with special reference to domestic animals. XVI. The influence of temperature and breeding upon the rate of growth of chick embryos. KAREL W. HENDRICKSON. Missouri Agr. Expt. Sta., *Res. Bull.* 140, 5-47(1930).—Temp. exerts a profound influence on the growth rate of chick embryos as measured by daily dry wt. and N content as indicated by temp. coefficients. This influence is greatest in the earliest stages of growth and it decreases as successive stages and practically ceases when the embryo is about 16 days old or when it reaches a dry wt. of approx. 2 g. Loss of moisture of eggs is slightly increased by an increase in temp. if the relative humidity is const. In general, the limits of the growth periods at 38.5° coincide with the periods marking a change in the chem. comp. utilized for energy. C. R. FRIEDMAN.

Respiration of the tissues. V. Respiration of blood, in vitro, of various homo-thermal animals. J. RICHIE AND R. S. SOBEL. *Arch. intern. physiol.* 31, 419-27(1929); *Physiol. Abstracts* 15, 220-1, cf. C. A. 21, 2109.—With the same content of erythrocytes the blood of animals with nucleated red cells has a greater respiratory intensity than that of animals with non-nucleated cells. There is no direct relationship between the oxyhemoglobin content of the blood and the O consumption. If the plasma is replaced by Ringer's soln., the respiratory intensity is diminished. G. H.

Muscular work and glucemia. V. Livera. *Arch. intern. physiol.* 31, 428-32(1929); *Physiol. Abstracts* 15, 248-B.—In the rabbit, starvation causes an immediate fall in blood sugar, followed by a const. level, intense work after several days' starvation produces a marked but very transient increase of the order of 100 mg. per l. G. H.

A new function of the stomach—its role in the metabolism of urea. R. V. AMESON, D. SIMES AND M. LYNCH. *Comp. rend.* 192, 308-10(1931), cf. C. A. 24, 5811.—A 14 kg. dog was anesthetized with chloralose and gastric secretion stimulated by injection of 2 mg. histamine. A sample of the gastric juice showed some NH_3 present but no urea, 45 min. and 90 min. after the intravenous injection of 15 g. urea. The gastric juice contained 1.5 and 2.7 times the normal amt. of NH_3 but no urea. When a 3% soln. of urea was kept at 37° in vitro in contact with washed dog stomach mucosa the urea was almost completely converted into NH_3 in a few hrs. Conclusion.—The normal stomach removes a little urea from the blood and excretes it as NH_3 . In times of renal insufficiency this action may be increased 5-10 times in amt. L. H. DUNBAR.

The relation between blood sugar and blood-clotting time. V. Inter-relation between blood sugar, clotting time and glycolysis. A. PARTOW. *Arch. ges. Physiol. (Pflügers)* 226, 317-33(1930), cf. C. A. 24, 4547.—The hyperglycemia following the administration of adrenaline or morphine is accompanied by an accelerated glycolysis. Insulin administration is accompanied by an inhibited glycolysis. No relation was found to exist between the rate of glycolysis and the clotting time of blood. VI. The relation of the blood sugar to the lactic acid content during glycolysis. A. PARTOW AND A. HERRON. *Ibid.* 407-11.—The lactic acid content of blood was always increased at the beginning of glycolysis. No relationship was discovered to exist between the sugar and lactic acid contents of the blood. ARTHUR GROLLMAN.

The constancy of the alveolar ventilation and the variations in the alveolar carbon dioxide and oxygen deficit during voluntary variations of the frequency and depth of respiration. HANS HUCKENRIE, H. FARRARSON, AND ERIC MOUNSEY. *Arch. ges. Physiol. (Pflügers)* 226, 414-30(1930). ARTHUR GROLLMAN.

The oxygen deficit and carbon dioxide tension of the alveolar air. HANS HECX-SCHTER *Arch ges Physiol (Pflügers)* 225, 431-47(1930), cf C A 24, 4750

ARTHUR GROLLMAN

The chemical composition of the lung: a contribution to the metabolic physiology of the lung. H. MISCHNAY, H. SCHÖNBERG AND P. JENKERSOHN *Arch ges Physiol (Pflügers)* 225, 445-63(1930) —Data are given for the H₂O, glycogen, fat, total N, total i. lypoidal i and Ca contents of dog lungs. The H₂O, glycogen and fat contents of the liver, muscle and heart of the same animals were also detd. The effect of the type of food ingested, glycogen feeding, starvation, phlorizin, insulin and adrenaline myoclonia, on the chem. compn. of the above mentioned organs are also given. A. G.

Investigations of the chemical composition of the various lobes of the liver. W. SCHLÖTTER *Arch ges Physiol (Pflügers)* 225, 481-90(1931) —The glycogen, fat, total N, ret N and H₂O contents of the various lobes of the liver differ to an appreciable extent. Various parts of the same lobe also differ as regards their chem. compn. A. G.

The heat formation of skeletal muscle when lactic acid formation is inhibited. ERNST FUCHS *Arch ges Physiol (Pflügers)* 225, 550-17(1931) —The caloric quotient of frog muscles poisoned by CH3BrCOOH is not altered from the normal. The poisoned muscles fail to show the evolution of an aerobic recovery heat. A. G.

The complex nature of the hypophyseal sex hormone: its action on the female. E. DÖRGER AND S. E. DE JONGH *Arch ges Physiol (Pflügers)* 225, 543-6(1931) —The sex hormone of the hypophysis consists of at least 2 substances. One factor causes maturation of the Graafian follicles, the other causes lacta formation. A. G.

The action of the hypophyseal sex hormone on the male organism. S. E. DE JONGH *Arch ges Physiol (Pflügers)* 225, 547-59(1931) —Young male rats or mice were treated with exts. of *mercurialis perennis*. Such exts. contain only the hypophyseal sex hormone. The physical action of this hormone was demonstrated.

ARTHUR GROLLMAN

The absorption of dyes by the red corpuscles. R. HÖDER AND G. PEYRILL *Arch ges Physiol (Pflügers)* 225, 555-59(1931) —The distribution of dyes between red blood corpuscles and serum is detd. by the nature of the dyes—particularly its soly. relationships and by the lypoid content of the corpuscle. A. G.

ARTHUR GROLLMAN

The action of potassium and calcium ions on the function of the isolated salivary gland supplied with d-generated nerves. O. V. NIKOLAY *Arch ges Physiol (Pflügers)* 225, 653-93(1931) —If the [Ca⁺⁺] of the perfusate through an isolated salivary gland, whose nerve supply is d-generated, is increased, there is an increase in the salivary secretion. An increase in the [K⁺] diminishes the secretion. In both cases there is a diminished blood flow through the gland due to vascular constriction.

ARTHUR GROLLMAN

Sleep, sleeplessness and physical exercise and their effect on the blood chemistry. II. The recovery period. H. HÖCHBERG, J. MICHELSEN AND H. BECKER *Arch ges Physiol (Pflügers)* 225, 738-47(1931) —The blood chemistry (lactic acid, O capacity, and p_H) of individuals deprived of sleep differs in the recovery period after exercise, from that of normally rested individuals.

ARTHUR GROLLMAN

The humoral nature of nervous irritability. I. P. RAFFENOV AND A. N. POBELINA *Arch ges Physiol (Pflügers)* 225, 767-93(1931) —Stimulation of the vagus leads to the formation of substances in the glandular cells of the stomach and pancreas which act as hormones in the circulation.

ARTHUR GROLLMAN

The relation between the hormone of the anterior lobe of the hypophysis (prolin) and the male sex gland. M. BERRY *Dent med Wochschr* 56, 1117-20(1930) —The injection of prolin, an ext. of the anterior lobe of the hypophysis, into young male mice caused marked degenerative changes in the testicle. In the case of older mice, the effect was less pronounced.

ARTHUR GROLLMAN

The importance of the thiocyanate present in gastric juice. G. LOCKERMAN AND W. LUTHER *Dent med Wochschr* 56, 1199-2(1930) —The presence of SCN⁻ in gastric juice greatly enhances its bactericidal power.

ARTHUR GROLLMAN

Hormones. BENJAMIN HARKOW *J Chem Education* 8, 621-9(1931) —A review

E. H.

The coenzyme of lactic acid formation of the muscle. KARL LOHMANN *Nachrichten* 19, 150(1931) —Previously the coenzyme of lactic acid formation in the muscle was found to consist of an autolyzable and a non-autolyzable part, the former being adenylypyrophosphoric acid (C. A. 24, 145, Meyer, C. A. 25, 2566). It was now found that the non-autolyzable part can be replaced by Mg salt. The system dialyzed muscle ext. + mag. phosphate + adenylypyrophosphate + Mg salt is able to split

glycogen into lactic acid. Absence of any of the 4 components inactivates the mixt. Muscle ext. inactivated after prolonged standing (4 days at 0°) by enzymic demolition of the adenylypyrophosphate (Meyerhof, *C. A.* 20, 1635) can be reactivated by addn. of this substance or of adenylic acid but to a lesser extent. A slight reactivation is caused by inosinpyrophosphate acid, no action is found on addn. of inosinic acid. An adenylypyrophosphate acid similar to that in the skeleton muscles of vertebrates or that of non-vertebrates occurs in heart muscle. The Harden Young hexosediphosphoric acid and Limbden hexosemonophosphoric acid are only cleaved by adenylypyrophosphate in the presence of Mg, likewise for the reverse reaction. An additional case where this combination is essential to effective action is in alc. fermentation by dialyzed maceration juice or by washed yeast (washed in weakly acid soln.). The system $Mg + \text{adenylypyrophosphate}$ fulfills all conditions for a coenzyme. Its relation to that of Euler and Myrbäck (*Z. phys. u. Chem.* 1923-30; cf. *C. A.* 24, 5312) is uncertain. The system is not active for methylglyoxalase (Neuberg, *C. A.* 7, 2115, Dakin and Dudley, *C. A.* 7, 1532).

B. J. C. VAN DER HOEVEN

The surface potential of the gastric mucous membrane in warm-blooded animals under various conditions. KURT KRAMER. *Z. Biol.* 91, 26-36 (1931).—Surface potentials of the gastric mucous membrane were derived by using KCl, HCl and NaOH in different concns. The direction and intensity depend entirely and characteristically on the soln. It seems that HCl can vary the permeability if its concn. is varied. Contrary to other membranes, these changes are very small and reversible. NaOH causes irreversible changes.

FRANCES KRAEHW

Investigation on the fluctuations of the potential of the gastric mucous membrane in warm-blooded animals. HANS SARRK. *Z. Biol.* 91, 37-42 (1931).—The initial potential is varied by changing the soln., in contact with the membrane. When acid gastric juice is in contact with the membrane, the potential is always negative, of variable height and depends upon the condition of the animal. At death the potential falls steadily and reaches positive values.

FRANCES KRAEHW

The influence of oxygen on the generation of lactic acid in the muscle of warm-blooded animals. ARNOLD HAYN, E. FRIEDRICH AND H. NIEMER. *Z. Biol.* 91, 53-62 (1931).—In work on pig muscle it was found that in the presence of O_2 the lactic acid content of muscle mash may be kept const. The amt. of total carbohydrate remains const. or decreases very slightly.

FRANCES KRAEHW

The activity of the heart and the central nervous system of mammals during great oxygen deficiency. L. ASNER AND N. SCHNITZKE. *Z. Biol.* 91, 66-72 (1931).—In certain cases it was possible to maintain normal breathing and circulation by inspiring air contg. only 3% O_2 .

FRANCES KRAEHW

Hydrogen-ion concentration of the alimentary tracts of fowl, cat and rabbit. ALVAN R. McLAUGHLIN. *Science* 73, 191-2 (1931).—Fowls, rabbits and cats were killed very rapidly and the pH value of the contents of several portions of the alimentary canal was measured by means of a quinhydrone electrode. The fowls and rabbits were killed very soon after feeding. All portions of the alimentary canal of the former were acid, while those of the latter were slightly alkaline excepting the stomach (pH 1.83) and cecum (pH 6.26). The results for the cats were uncertain because of the lack of knowledge of their past history, all portions were definitely acid.

I. M. LEVINE

The alleged presence of "bound potassium" in muscle. WM. E. CALLISON. *J. Biol. Chem.* 93, 165-8 (1931).—The decrease in "bound K" in muscle tissue placed in isotonic saline is related to the amt. and frequency of change of isotonic NaCl soln. and to the size of the pieces of muscle tissue. Autolysis for a period of 21 hrs. did not apparently affect the diffusion of K into the soln. The present expts. furnish no evidence for the existence of "bound K" in muscle tissue" as claimed by Neuschloss (*C. A.* 20, 5497, 5498) and corroborate the findings of Raab (*C. A.* 21, 3227) and Hober (*C. A.* 23, 3745). Neuschloss' method for the detn. of "bound K" is unsatisfactory.

A. P. L.

Studies on arginine. I. The rate of catabolism of arginine in rats, including a method for the determination of arginine in biological material. VROO C. KISCH, JAMES M. LUCK AND ANNA EVELYN SMITH. *J. Biol. Chem.* 90, 677-96 (1931).—A volumetric method, in principle resembling that of Hanot and Cahn (*C. A.* 21, 1471), is described for the detn. of very small amts. of arginine in protein free tissue exts. The proteins are pptd. by H_2WO_4 , the arginine is converted by arginase into ornithine and urea and the urea is detd. as diacetylhydrol urea by oxidation with $K_2Cr_2O_7$ and H_2SO_4 , according to the vol. procedure of Allen and Luck (*C. A.* 23, 3206). Muscle tissue of fasting female rats contains an av. of 26.1 mg. of arginine per 100 g. and the whole carcass 27.5 mg. If administered subcutaneously, arginine is rapidly catabolized and the ornithine which would presumably be formed must be catabolized almost as rapidly

as less than 12% of the injected N is present at any time in substances other than urea and arginine. Bibliography of 28 references. A. P. LOTNAOF

The distribution of blood sugar. MICHAEL SOMOGYI. *J Biol Chem* 90, 731-5 (1931). —Folin and Svedberg (*C A* 25, 335) have recently reported a ratio of 60:100 for the distribution of corpuscle and plasma sugar for healthy persons and claim that the corpuscles in the blood of diabetics contain relatively more sugar than those of non-diabetics. S. has redetd the values for true sugar using his new method of deproteinization (*C A* 25, 2165) and obtains results approx. the same as those previously reported (*C A* 22, 3204), 77:100, in both diabetics and non-diabetics. F. and S. detd total reduction values and not true sugar in plasma and slight and negligible as the amt. of non sugar reducing substances in plasma may appear (5-6 mg %) it suffices grossly to distort the true picture of the distribution of sugar. If the figures are corrected to true sugar values by the deduction of 5 mg % of reducing non-sugars, a rise in the distribution ratios occurs and at once removes the apparent difference between the ratios of normal and diabetic individuals which is actually non-existent. The corrected figures are well in accord with those obtained by S. A. P. LOTNAOF

Local metabolism and tissue reaction. Further investigations of the relations between local metabolism and cell growth. GERHARD MAYRA-DÖRKIN. *Beitr. path. Anat.* 85, 555-64 (1930). —The growth-stimulating effect of solns. of inorg. salts upon the endothelial cells of the lymph spaces of white mice observed in earlier expts. (*C A* 24, 5358) was not observed in later expts. carried out under more carefully controlled conditions. F. R. MARV

The physiologic variations in the inorganic blood phosphorus content at the different age periods. An attempt to explain these in the growing child. J. K. BULLOCK. *Am. J. Diseases Children* 40, 725-40 (1930). —The inorg. P content of the serum of newborn infants is 4.82 mg per 100 cc. It increases rapidly to a value of 5.63 mg per 100 cc. at the 7th month. From this time it tends to diminish gradually until the 19th or 20th yr., the period of termination of osseous growth, and then remains more or less const. The age periods of most rapid growth do not show as marked an increase in inorg. P as the seasonal periods of most rapid growth. E. R. MARV

The intermediate metabolism of fats. The influence of the administration of certain fatty acids of the aliphatic series on ketogenesis and on the acid-base equilibrium of the organism in diabetics and on patients suffering with liver disturbances. L. CANNAVO. *Boll. soc. ital. biol. sper.* 5, 805-10 (1930). —The acidosis caused by the administration of various fatty acids and the relation of acidosis and the production and elimination of "ketone bodies" were studied clinically on patients. From the data accumulated, it appears that all or nearly all the acids of the aliphatic series produce acidosis, however, only those which contain an odd no. of C atoms are intensely ketogenic. The acidosis caused by the administration of acids with the odd no. of C atoms may be due to non ketogenic org. acids formed from the former by the degradation of their mols. Whenever, as a result of the administration of fatty acids there arose a more or less intense acidosis, an increase of NH_3 in the urine was observed. In patients with hepatic lesions a fair amt. of ketones was observed in the blood and urine, this points to a diminution of the function of the hepatic cell which in all probability is a factor in hepatic acidosis and coma. PETER MASUCCI

The reticulo-endothelial system in the defense of the organism. LAMBERTO LENZI. *Boll. soc. ital. biol. sper.* 5, 825-8 (1930). —Contrary to what has been observed with inorg. poisons, animals which have been previously stimulated with metallic colloids (collargol) do not become more resistant to nrg. poisons (morphine HCl or quinine sulfate). PETER MASUCCI

Experimental studies on the influence which blood plasma exerts on arterial pressure in contrast to that exerted by the corpuscles. LUIGI SPADAFINA. *Boll. soc. ital. biol. sper.* 5, 829-32 (1930). —Repeated infusions of distd. H_2O in dogs such that the no. of corpuscles was reduced to $\frac{1}{2}$, or even $\frac{1}{4}$ of the normal did not appreciably change the arterial blood pressure. Injections of pyridine, subcutaneously or intravenously, reduced the corpuscles to $\frac{1}{2}$ the normal, but did not change the arterial pressure. These results lead S. to conclude that plasma plays a much more important role in the maintenance of the normal blood pressure than the blood corpuscles. PETER MASUCCI

The sedimentation of the blood studied by means of the nephelometer. IV. The influence of the viscosity of the medium. A. AGGAZZOTTI AND G. BUCCIARDI. *Boll. soc. ital. biol. sper.* 5, 867-9 (1930). cf. *C A* 24, 1897. —By increasing the viscosity of Ringer soln. by the addn. to it of glycerol, starch or glucose, and using the modified Ringer as a diluent for blood (9.1% citrated blood), the variations in the luminosity which precede and follow sedimentation of the corpuscles are changed markedly. The

increase of the initial luminosity which precedes the beginning of sedimentation is noticeably retarded.

The permeability coefficient of aqueous solutions of lactic acid on muscle. P. ROWINSKI. *Bull. soc. sci. nat. ser. 5*, 899-901(1930)—Frog muscle preps. were perfused with Ringer soln. (without bicarbonate) and after having frozen and powdered the muscles with liquid air, lactic acid was detd. (1) *Perfusion with Ringer soln. contg. lactate*—The lactic acid in the muscle increased with increasing concn. of the salt in the perfusion liquid. (2) *Perfusion with Ringer soln. contg. lactic acid*—The concn. of lactic acid in the muscle varied according to the concn. of the acid in the perfusion liquid. Up to 40 mg % (fn 3.0) the lactic acid in the muscle increased rapidly, but with greater concns. the increase became slower. (3) *Perfusion with Ringer soln. contg. HCl*—When the liquid came out of the muscle had a fn of 6.8 the amt. of lactic acid was normal, but when the muscle failed to regulate the fn of the circulating liquid, the amt. doubled.

P. M.

Studies on the surviving hepato-pancreatic preparation. I. The metabolism of fats. OSCAR CANTONI. *Bull. soc. sci. nat. ser. 5*, 918-21(1930)—The fatty acid content of the blood did not vary noticeably in fasting animals but did decrease in dogs with alimentary lipemia. The circulating fatty acids apparently went through a process of desaturation as evidenced by the lowering of the m.p. and solidification point, and by the increase of the iodine no. II. The metabolism of cholesterol. *Ibid.* 921-2. The diminution of cholesterol during circulation was gradual, const. and independent of the amt. of alimentary fat. III. The respiratory exchange. *Ibid.* 922-4. The R.Q. was generally low, indicating intensive oxidative processes. IV. The uropoietic function of the liver. *Ibid.* 924-5. There was an increase of urea in the blood flowing from the liver. The increase was gradual, const. and bore no relation to the alimentary fat.

The lipoidal hormone of the hypophysis and the vitamin of fertility. RENZO AGNOLI. *Bull. soc. sci. nat. ser. 5*, 931-9(1930)—Some preliminary evidence is offered that the lipoidal hormone of the anterior pituitary lobe is similar to the fat sol. vitamin E.

The action of adrenaline on the reduction processes of frog muscle pulp. M. COMEL. *Bull. soc. sci. nat. ser. 5*, 932-4(1930)—With methylene blue or *m*-dinitrobenzene as indicators, adrenaline in concns. of 10^{-4} to 10^{-11} when added to frog muscle pulp did not accelerate or inhibit reduction processes.

Experimental uriccolysis in the heart-lung preparation. GINO BERGAMI. *Bull. soc. sci. nat. ser. 5*, 934-6(1930)—Uric acid injected into the heart lung prep. is not destroyed. The addn. of uricase to the blood circulating in the prep. causes only a slight diminution of the uric acid content.

Parathyroids and irradiated ergosterol. G. GANTINI. *Bull. soc. sci. nat. ser. 5*, 953-5(1930)—Irradiated ergosterol cannot be substituted for the parathyroid hormone, the former merely retards tetany without influencing cachexia.

The influence of nonelectrolytes on the buffer action of animal and vegetable organs. GIUSEPPE ROSSO. *Bull. soc. sci. nat. ser. 5*, 953-5(1930)—The aim was to establish whether a given organ displaces with equal intensity the reaction of 2 isotonic solns. of equal fn. The nonelectrolytes used were urea, glucose, sucrose and galactose; the electrolyte was NaCl. The results show that nonelectrolytes aid in an acid reaction and inhibit in an alk. reaction the unfolding of the buffering action of tissues, e.g., NaCl and glucose soln. had initial fn. of 5, after shaking both with equal amts. of frog leg tissue and for equal lengths of time, the NaCl had a fn. of 7.18, the glucose fn. 7.45. When the initial fn. was 5, after the solns. were shaken with the tissue, the NaCl soln. had a fn. of 5.65, the glucose fn. 5.93. This behavior explains in all probability the action of nonelectrolytes on the permeability of protoplasm.

The so-called influence of the suprarenal medulla on alimentary hyperglucemia. S. RACCHUSA AND A. CRIVARA. *Bull. soc. sci. nat. ser. 5*, 993-6(1930)—Dogs subjected to the complete or incomplete ablation of the suprarenal capsule showed, immediately after the operation, a marked lowering of temp. which reached 34-35° in some cases. After the complete ablation of the suprarenal medulla, and in some cases after the removal of the greater part of the cortex, not only was there no hypoglycemia but in some cases there was a tendency toward hyperglucemia after the administration of glucose.

The direct action of follicular liquid on the motility of the tubes. LUIGI MANZI. *Bull. soc. sci. nat. ser. 5*, 1020-5(1930).

The energy of movement of spermatozoa. E. E. IVANOV. *Z. Tierpsych., Zuckersgshw. Tierarzht.* 20, 404-18(1931)—The onward movement of dog spermatozoa

group. Ammoniacal AgOH colors the yellow rods an intense black, demonstrating a reducing power, this is not due to an aldehyde group, however, because decolorized fuchsin does not have its color restored, neither is it due to the presence of 2 phenol groups, nor to the presence of one phenol and one aromatic amine group, since the chromaffin reaction is negative, on the contrary, prolonged oxidation by alk dichromate causes the rods to disappear, an irreversible reaction as shown by treatment with reducing agents. The experiments of Bach and Haehn on the action of tyrosinase on tyrosine make it possible to deduce the presence of a quinonoid nucleus, *in vitro* the quinone is condensed very easily with tyrosine to give deriva with properties very similar to those described for the melanin pigment, therefore I. concludes that this latter substance is a condensation product of a compd having a phenolic base like tyrosine, with a deriv. possessing a quinonoid structure. This compd, first stage of melanogenesis, preserves intact a phenolic hydroxyl.

ALBERT I. RAWLINS

The relation existing between the factors determining the amount of oxygen consumption of the organism. G. LITAVSKY, H. AUGUST, I. COMTESSCO AND B. NIKOLSKY. *Compt rend soc Biol* 106, 108-10 (1931).—The following equation is derived

$$O_2 = \text{Output} \times Hb/O_2 \div \left\{ \frac{[100x_a^2]}{(1/k_a) + x_a^2} - \frac{[100x_v^2]}{(1/k_v) + x_v^2} \right\}, \text{ where}$$

output is the quantity of blood which passes through the capillaries from the arterial to the venous system, expressed numerically by considering 100 cc. as unity, Hb/O_2 is the total O_2 capacity of the blood, $1/k_a$ and $1/k_v$ are the constants of dissociation of the oxy-hemoglobin of arterial and venous blood, resp., x_a and x_v are the tensions, resp., of arterial and venous O_2 expressed in mm. of Hg.

ALBERT I. RAWLINS

The growth-promoting hormone of the pituitary body. H. B. VAN DYKE AND ZERJA WATSON LAWRENCE. *J. Pharmacol* 40, 413-22 (1930).—An alk. ext. of the anterior lobe, which appears to be free of the luteinizing principle, further purified by filtration through a Berkefeld filter, pptn. with Na_2SO_4 , dialysis, etc., is described. A method of assay is given; it involves the use of hypophysectomized or normal adult rats. C. R.

A critical study of the assay of female sex hormone preparations. P. R. D'AMOUR AND R. G. GUSTAVSON. *J. Pharmacol* 40, 473-84 (1930).—The work of Coward and Burn (C. A. 22, 632) on standardization of female sex hormone prepn. was repeated. D'A. and G. agree with Coward and Burn that the rat unit should be the amount of hormone necessary to produce estrus in 50% of a large group of animals, and that a standard unit be used by all workers.

C. RINGEL

The preparation and assay of crystalline female sex hormone. P. R. D'AMOUR AND R. G. GUSTAVSON. *J. Pharmacol* 40, 485-9 (1930).—Prepn. of a cryst. extrin is described which on assay showed 1 rat unit equal to 0.284.

C. RINGEL

The diuretic hormone of the brain. J. OLIVIER. *Munch med Wochtschr*, 77, 18-9 (1930).—A substance with a strong diuretic action is mobilized by stimulations of certain portions of the brain. The substance was isolated to the extent that it was active when administered to other animals. The hormone is formed in the midbrain.

R. C. WILLSON

NEWBURN, LOUIS H., AND JOHNSON, MARGARET W. The Exchange of Energy between Man and the Environment. Springfield, Ill.: C. C. Thomas. 104 pp.

G—PATHOLOGY

H. GUDON WELLS

Presence of pancreatic enzymes in the bile. H. L. POTT. *Werner klin. Wochtschr*, 42, 800-7 (1929).—In certain pathol. conditions the bile, which usually contains no diastatic enzyme, showed high diastatic values.

B. C. A.

Ketosis in seasickness. J. R. MARRACK. *Brit. Med J.* 1931, I, 178-9.—Ketosis occurs in seasickness after vomiting appears, sometimes before. Ketosis is not the cause of seasickness, but rather is an evidence of some unknown metabolic disturbance.

J. B. BROWN

Relation of calcium to the hemorrhagic tendency in obstructive jaundice. ROBERT R. LINTON. *Ann. Surgery* 93, 707-21 (1931).—The lowered coagulability of the blood in cases of obstructive jaundice does not seem to be due to a deficiency or abnormality of the blood Ca. The most effective means of preventing and stopping post-operative hemorrhage in obstructive jaundice is repeated direct blood transfusions used in conjunction with a high carbohydrate and fluid intake.

RACHEL BROWN

is not a deficiency. However, a deficiency may be one of several possible causes of gitter.

H. R. MAIN

"Uremia" following protracted vomiting. F. A. ALBANI. *Klin. Wochschr.* 10, 1557 (1933). Protracted vomiting may cause azotemia, increased nonprotein nitrogen in the blood, anuria and even convulsions, without any anatomical changes in the kidneys. In one case, the administration of insulin, glucose and physed. NaCl solution caused the disappearance of all symptoms.

H. R. MAIN

The color of the urine in diffuse hematogenous kidney disease. RAVIN. *Biochem. Klin. Wochschr.* 10, 1573 (1933).

H. R. MAIN

The demonstration of brain antibodies in human syphilis. F. GRIENHARDT and G. FRECHER. *Klin. Wochschr.* 10, 207 (1933), cf. *C. A.* 24, 4462. The mere fact that an ale. ext. of brain may give complement fixation with syphilitic spinal fluid and not with blood from the same patient does not prove the presence of brain antibodies in the fluid.

The same phenomenon can be produced with any organ ext. by proper diln. and is due to the fact that spinal fluid continues to react even with high dilns. of ale. ext., while the serum does not. A mixed brain ext. gives complement fixation with the spinal fluid and serum from many non-syphilitic, organic neurological conditions, which strongly indicates that there may be a true antibody to brain fluids.

H. R. MAIN

The electrolyte content of the blood in a case of diabetes insipidus. M. CHATWANT. *Deut. med. Wochschr.* 58, 1700 (1933). The blood in a case of diabetes insipidus showed an increased K, a decreased Ca, and an increased Na content. The H₂O binding capacity of the corpuscles was increased.

ARTHUR CHISHAM

Value and duration of immunity conferred by antiphosphatidase anatoxin. Titration of antitoxin present in the serum of vaccinated children. G. RAMIN and ROBERT DRENT. *Ann. J. Diseases Children* 41, 17 (1933), cf. *C. A.* 24, 2171. The antitoxin content of the blood serum of children who have a neg. 1 c.c. test after immunization with phosphatidase anatoxin is more than $\frac{1}{100}$ unit per c.c. It is apparent the same whether 2 or 3 injections of anatoxin are given but may be higher if a 4th injection is given and if there is a prolonged interval between 2 of the injections. It appears to be higher the older the child is when inoculated and does not tend to diminish with increase in the time after inoculation.

W. R. MAIN

Cell and plasma chloride in the pyloric stenosis of infants. MONTAGUE MAIZLIN and CATHERINE B. McARTHUR. *Ann. J. Diseases Children* 41, 35 (1933). The blood plasma of normal infants less than 12 months of age contains 2.2-2.8 millimoles of bicarbonate ion and 90-100 millimoles of chloride ion. The cells contain 11-50 millimoles of chloride ion. In pyloric stenosis in infants, the pH of the plasma and the content of bicarbonate are normal or increased. The chloride content of the plasma and of the cells is often decreased, the chloride content of the cells usually being lowered to a greater extent than that of the plasma. The vol. of the blood often is decreased and the relative vol. of the cells increased.

W. R. MAIN

Dietary protein and blood-clotting function. I. NEWTON KUHLMANN and IRMA LOUISE SAMUEL. *Ann. J. Diseases Children* 41, 18 (1933), cf. *C. A.* 24, 3517. Prolonged feeding of a diet rich in protein tends to maintain the blood clotting factors at a high level, thereby increasing the coagulability of the blood. An increase in the production of fibrinogen is induced by injury of the tissues through surgical intervention, and a transient increase by the introduction of heterologous proteins into the vascular system.

W. R. MAIN

Ossification. 1. Callus formation and calcification. I. NEWTON KUHLMANN and RICHARD N. BERG. *Ann. J. Diseases Children* 41, 26 (1933). Bone repair appears to be a local rather than a systemic process and may be, therefore, retarded or accelerated by locally applied agents. The degree and rate of calcification in experimentally induced fractures in rabbits are directly proportional to the amt. of available fibrous tissue. The amt. of the bones may be retarded by soln. of the fibrous tissue by means of the injection of an alk. soln. of trypsin at the site of the fracture. It may be accelerated by the stimulation of the production of fibrous tissue induced by the local injection of tissue extract. The local production of fibrous tissue does not appear to be altered by a diet high in protein. The Ca content of the blood tends to be higher and the phosphate content lower in the course of normal bone repair. A more rapid rate of calcification is associated with more marked alterations in the Ca and phosphate contents.

W. R. MAIN

The proteolytic activity of the spleen in disease and hemorrhage. ARNOLD STRAUSS. *Deut. path. Anat.* 85, 251 (1933). The softening of the pulp tissue of spleens obtained at autopsy is associated with an increased proteolytic activity, which may be measured by the ability to liquefy gelatin or egg white. The optimum activity of the enzymes occurs

at pH 8 and pH 5. The enzymic activity is increased with the increase in time after death and is markedly increased in acute septic poisoning and post hemorrhagic swelling. It is normal in spleens with chronic septic poisoning, in indurated and atrophied spleens and in those of children under 3-5 yrs. The presence of leucocytes appears to be an important factor in the degree of enzyme activity, since there is no proteolytic activity or softening of the pulp in animal spleens which contain no leucocytic enzymes. The softening of the adrenals appears to be associated with a purely autolytic rather than a proteolytic process. E. R. MAIN

The reactions of tissues and cells at artificial boundaries in the animal body. K. AKAMATSU. *Beitr. path. Anat.* 85, 249-70 (1930).—The relation between the surfaces of small blocks of gelatin placed in the abdominal cavities of guinea pigs and the surrounding tissues appears to depend upon the concn. of H and OH ions within the blocks. The acid blocks have a marked tendency to adhere to and unite with the surrounding tissues. The alk. blocks have no tendency to unite with the tissues. The accumulation of leucocytes at the boundary surfaces and in the interiors is much more marked in the acid blocks. The leucocytes in the interior of the alk. blocks, although fewer in no., give a stronger test for the presence of oxidase. E. R. MAIN

The biologic and pathologic significance of copper, zinc and manganese. WALTER HERKEL. *Beitr. path. Anat.* 85, 513-51 (1930).—The continued feeding of various salts of Cu to rabbits and rats does not induce cirrhosis of the liver. Large amts. of the ingested Cu accumulate in the liver, small quantities in the kidneys and very little in the bones. The amt. of Cu stored depends upon the amt. fed and the duration of the period of feeding and not upon the form in which it is fed. The normal liver contains from 6.5 to 36.1 mg. of Cu per kg. of dry substance. The values are doubled in pregnancy and are many times greater in the new born. The Cu content of the liver in hemochromatosis is 24-384 mg. per kg. and is greatly increased in all other organs except the kidneys and bones. In cirrhosis without pigmentation the liver contains 67.6-374.9 mg. Cu per kg. The Zn content of the normal liver is 50-317 mg. per kg. of dry substance and is not increased in hemochromatosis. L. R. MAIN

The chemistry of hemosiderin. KURT W. SCHULTZE. *Beitr. path. Anat.* 86, 101-2 (1931).—The detn. of the amts. of ferric and ferrous ions present in the organs obtained at autopsy is extremely difficult, since the ferric ion is reduced rapidly during the 1st 6 hrs. after death and the ferrous ion is then gradually oxidized. However, if an appreciable amt. of Fe is present, as in hemosiderosis of the liver, the quantity of free ferric Fe in the liver is much greater than in other organs examd. under identical conditions. HClO and other preservatives of tissues hasten the oxidation of the ferrous ion. Expts. in which hemosiderosis is induced in animals by the injection of hemoglobin and dust H_2O , while not conclusive, indicate that a large part of the Fe of hemosiderin exists in the ferric state. Procedures are described for the extn. of the free Fe of tissues by dil acids with the exclusion of air and for the detn. of the total free Fe and the free ferrous and ferric Fe. E. R. MAIN

The Hinton glycerol cholesterol reaction for syphilis. Second modification. Wm A. HINTON AND ARTHUR BERK. *New Engl. J. Med.* 202, 1054-9 (1930).—Supplementary procedures for the glycerol cholesterol reaction previously proposed (*C. A.* 24, 3828) are described which involve the use of a cholesterolized beef heart ext. and the formation of a characteristic ring instead of a ppt. E. R. MAIN

Blood cholesterol values in hyperthyroidism and hypothyroidism. Their significance. ROBERT L. MASON, HAZEL M. HUNT AND LEWIS HURXTHAL. *New Engl. J. Med.* 203, 1273-8 (1930).—The cholesterol content of the blood tends to become decreased in hyperthyroidism, the lowest values being associated with severe toxemia. It becomes markedly increased in true myxedema. E. R. MAIN

Studies on anaphylaxis. VIII. The influence exerted by the liver on anaphylactic ictus and on peptone ictus, with particular reference to the experimental method and sources of error. ALBERTO MARRASSINI. *Boll. soc. ital. biol. sper.* 5, 833-7 (1930).—Guinea pigs were sensitized to egg albumin and to Witte peptone. The test dose was given after 19 days, part of the pigs received this dose in the jugular vein and part in a branch of the mesenteric vein. Amts. of egg albumin below certain limits when used for the test dose gave very different results with animals of the same weight. The shock varied from 0 to 6 according to the Arlong and Wauthy scale. The amt. of albumin which might produce a mortal ictus or an ictus of varying intensity even to total absence was about $\frac{1}{2}$ the dose which produced death with certainty in 250-300-g. guinea pigs. In guinea pigs 250-350 g. or over there was no relation between wt. of the animal and the dose of antigen necessary to produce a shock of equal intensity. These facts are ascribed to profound individual differences. 0.00150 g. albumin almost always

produced death; 0.00175–0.00250 g constantly produced death. The results of slight or no shock reported by many authors who injected larger amts of albumin must be accepted with much reserve, the same applies to results obtained with horse serum in doses of 0.05 cc, which corresponds to 0.003625 g protein. The results obtained by injecting the test dose into the portal circulation were similar to those obtained when injected into the jugular vein. No definite ratio was found between the quantity of peptone necessary to produce sure death and the wt of the animal. The weakest dose of peptone which produced a sure violent shock in 250–350 g guinea pigs when injected into the jugular vein, when injected into the portal circulation gave slight or no shock in about $\frac{1}{4}$ the animals. PETER MASUCCI

Alkaline reserve in septic exsces. RODOLFO REDI. *Boll soc ital biol sper* 5, 852(1930).—The alk. reserve values could not be used as a prognosis either for the duration or course of the disease. PETER MASUCCI

Studies on the acid-base equilibrium of the blood of septic (alkaline reserve, hydremia, glucemia, chloremia and urological examinations). RODOLFO REDI. *Boll soc ital biol sper* 5, 853–4(1930).—The causes which change the acid base equil of the blood of septic patients were investigated. The attempt was made to see whether the changes were due to the formation of an excess of acid substances, to a diminution of their neutralization or to their elimination. The results lead R. to state that there are too many factors to take into account to arrive at any definite conclusion. A knowledge of the alk. reserve in septic cases is of value only when considered along with other clinical findings. PETER MASUCCI

The adsorption of the virus of chicken sarcoma by hemoglobin. CLAUDIO PULCHER. *Boll soc ital biol sper* 5, 888–91(1930).—Electropos. hemoglobin adsorbs the virus, electroneg. hemoglobin does not. These results are in accord with cataphoresis expts and show that the virus is negatively charged. PETER MASUCCI

The pH and alkaline reserve of the blood in experimental intestinal occlusion. M. MAIRANO AND G. VECCHI. *Boll soc ital biol sper* 5, 912–5(1930). cf. C A 24, 5832.—The exptl. intestinal occlusion obtained by sectioning the intestine provokes a progressive increase of the alk. reserve which runs parallel to the clinical picture. The increase of the alk. reserve bears no relation to the height of the occlusion. The pH of the blood increases slightly, reaching a max. of 7.50–7.55 at the terminal period of the occlusion. PETER MASUCCI

Uroroseinuria and calciuria as indices of hepatic insufficiency. S. GULLOTTA. *Boll soc ital biol sper* 5, 971–3(1930). PETER MASUCCI

The acid-base equilibrium in dogs after cutting the vagi. G. QUAGLIARIELLO AND E. TRIA. *Boll. soc. ital. biol. sper.* 5, 1016–9(1930).—After cutting both vagus nerves, the changes in the acid base equil of the blood were slight, there was a slight acidosis but this remained within limits compatible with life. There was also a slight diminution of blood bicarbonates. Marked changes, however, were noted in the urine. The concn of bicarbonates increased and the acidity decreased because of the increase in NH_3 . PETER MASUCCI

The influence of protein substances on the production of antibodies. I. G. ASCIONE. *Boll soc. ital biol sper.* 5, 1038–41(1930).—Beef serum albumin injected into rabbits 48 hrs. after each injection of typhoid or staphylococcus vaccine does not exert any influence on the production of antityphoid and antistaphylococcal agglutinins. The globulin of the same serum injected similarly inhibits the production of antityphoid agglutinins but accelerates the production of antistaphylococcal agglutinins in a concn 3 times greater than the controls. II. *Ibid* 1042–4.—Two different strains of paratyphoid A were used. The rabbits immunized with bacteria and with lactalbumin yielded a serum with an agglutinating power no different from the controls. Similar results were obtained with rabbits immunized with bacteria and serum globulin. However, rabbits immunized with bacteria and serum albumin yielded a serum with a concn. of agglutinins 3 times greater than the animals receiving the bacteria alone. P. M.

A substance contained in lymph and lymphnodes producing hypoglycemia. PRO. MARFORI. *Arch. fisiol.* 28, 142–9(1930).—A fistula of the thoracic duct in a dog produces hyperglycemia but not glucosuria. Lymph and lymphoganglion compensate this increase of blood sugar. Both have a weak and inconst. action on the pancreas of the diabetic dog. The action is not due to insulin, but to a hormone antagonistic to adrenaline and acting by reduction of the activity of the vegetative sympathetic system. A. E. M.

The pituitary factor in arteriosclerosis. Its experimental production in rabbits. ROBERT C. MOHRILIC AND EUGENE A. OSTUS. *Ann. Internal Med.* 4, 579–92(1930).—Marked production of atheromatous plaques in the rabbit aorta may be produced within

100 days by the feeding of a diet high in cholesterol and the daily injection of posterior lobe pituitary ext

JOHN T. MYERS

Natural bactericidal antibodies: observations on the bactericidal mechanism of normal serum. T J MACKIE AND H M FINKELSTEIN *J. Hyg* 31, 35-55(1931) — Serum complement has no bactericidal action *per se*. An antibody like substance invariably acts as an intermediary agent. This sensitizing agent is stable at 55° but labile at 60° to 65°. It is resident mainly in the fraction of the serum which is insol in carbonic acid. It is present in the serum of young animals before certain other natural antibodies develop. The natural bactericidal antibodies are sp for certain bacteria. A non sp extracellular substance occurs in bacterial cultures which may neutralize or inhibit these antibodies. This substance is liberated in large amts when the culture is heated to 120°. It can be removed by repeated washing with saline soln. Strains of bacteria vary in their production of this substance. JOHN T. MYERS

The results of some quantitative experiments on the serum precipitation reaction. G L TAYLOR *J Hyg* 31, 56-83(1931) — Antigen and antibody react in equiv. proportions, in the serum pptn reaction, to form an antigen-antibody complex which is adsorbent for suitable material in the surrounding medium. Excess of antigen is adsorbed and prevents completion of the reaction. Excess of antibody has no effect. Speed of particulation depends on the amt of antigen antibody complex and adsorbable matter present. There is some evidence that in certain cases the nature and amt of adsorbable matter present may be such as to cause removal of excess antibody from the supernatant liquid. In such cases the adsorbable matter, not the antigen-antibody complex, may be adsorbent to the excess of antibody. The addn. of adsorbable matter as normal rabbit or guinea-pig serum is a useful method of hastening and facilitating the reading of reactions. JOHN T. MYERS

Observations on *Salmonella* agglutination and related phenomena. The influence of fresh normal serum on agglutination. P BRUCE WHITE *J. Path Bact* 34, 23-37(1931) — In concns from 1/20 upward normal rat and rabbit sera inhibit sp somatic agglutination of *Salmonella* bacilli with certain partial exception of coliform and proteus bacilli. Normal guinea pig serum has similar properties but is much more apt also to have an agglutinative effect (conglutination). Normal serum has little effect on flagella agglutination. The inhibiting property disappears rapidly at 55°. Killed cultures, particularly alc. treated ones, are more readily inhibited. Hemolytic complement is necessary for this inhibition, but another more stable factor must be present. Inhibition fails if the NaCl concn. is more or less than 0.95%. JOHN T. MYERS

The effect of variations in the media on the growth of normal and malignant tissues *in vitro*. F C PYBUR AND H T FAWCETT *J Path Bact* 34, 39-44(1931) — Preheating at 55° confers an inhibiting action on normal tissue growth which is not compensated by complement, but favors the growth of carcinoma. Pregnant serum has a stimulating effect on tissue growth. Embryonic and testis ext stimulates the growth of embryonic tissues, the amt of growth in 24 hrs being double that in control cultures. J. T. M.

Schwartzman's phenomenon of local skin reactivity to bacterial products. P. M. BURNER *J Path Bact* 34, 45-60(1931) — A substance can be obtained by washing young bacterial cultures with isotonic NaCl soln., which on intradermal injection in rabbits produces a transient modification of the locally infiltrated tissues so that when the same substance is injected intravenously about 24 hrs later an acute purpuric and necrotic reaction appears over the infiltrated area 2 or 3 hrs later. This substance closely resembles bacterial endotoxins. Formalin treated exts give typical reactions when used as a skin preparatory agent but will function as the intravenous provocative agent. JOHN T. MYERS

Observations on melanosis coli. M J STEWART AND ELLA M HICKMAN *J. Path Bact* 34, 61-72(1931) — Melanosis coli is due to the deposition of a melanin like pigment within large mononuclear cells in the mucous membrane of the intestine, limited usually to the large intestine and appendix and ranging in color from gray or buff through dark gray or brown toinky black. Metastasis of the pigment may occur into the submucosa and the mesenteric lymph glands. The most acceptable theory of the mode of formation of this pigment is that of Pick, *s e*, that aromatic protein disintegration products are absorbed from the colon and converted into melanin within the connective tissue cells by the action of a tyrosinase-like enzyme. JOHN T. MYERS

The intermediate zone phenomenon encountered in certain Br. abortus agglutinating sera. F. W. PRIESTLEY *J. Path Bact* 34, 81-9(1931) — The nature of the agent causing inhibition at certain zones is obscure. It is pptd from serum with the agglutinin in the globulin fraction and is removed by the homologous organism. An increase in NaCl concn. will remove the zone; this indicates that the zone is due to an in-

sufficient reduction of potential at that point and not to a decrease of cohesive force

JOHN T. MYERS

The hypochlorhydria of asthma in childhood. GEORGE W. BRAY. *Quart. J. Med.* 24, 181-97 (1931).—Asthmatic children, in general, show a markedly diminished power of secreting acid gastric juice in response to a test meal. Acid therapy should be combined with allergic treatment.

JOHN T. MYERS

The electrical charge of the virus of foot and mouth disease. IRMENGARD SIEBERT-MOPROW. *Zentr. Bakt. Parasitenk., I. Abt.*, 119, 12-7 (1930).—The virus carries a negative charge. It is not known whether this is an actual property of the virus or of some substance which acts as a carrier.

JOHN T. MYERS

Lipoid antibodies. IV. The origin and proof of syphilitic blood changes. FELIX KLOPSTOCK. *Zentr. Bakt. Parasitenk., I. Abt.*, 119, 78-84 (1930), cf. *C. A.* 22, 3453.—There are lipid compounds in the spirochetes which produce specific antibodies in rabbits and horses. The infected human patient acquires the power of combining with these substances during the early course of the disease so that antibodies may not appear till later.

JOHN T. MYERS

Hypoglycemic intoxication in trypanosomiasis, etc. KURT SCHERN. *Zentr. Bakt. Parasitenk., I. Abt.*, 119, 297-302 (1931).—In experimental animal trypanosomiasis, liver function is disturbed. No yeast-fermentable substance can be extracted from the liver. Only minimal quantities of sugar remain in the liver and blood.

JOHN T. MYERS

The influence of bacteriophage on the partial antigens of several microorganisms. MINORU TOYODA. *Zentr. Bakt. Parasitenk., I. Abt.*, 119, 350-3 (1931).—After solin by bacteriophage, the partial antigens remain in the lysate, their structure is not changed appreciably.

JOHN T. MYERS

Enzyme studies on the virus of foot and mouth disease. K. POPPE AND G. BUSCH. *Zentr. Bakt. Parasitenk., I. Abt.*, 119, 398-406 (1931).—The virus contained amylase, lipase and oxidase, but catalase, reductase and urease were not found. Virulence increases with the amount of lipase, oxidase and amylase. In infected guinea pigs there is an increase of blood amylase but not of catalase.

JOHN T. MYERS

The true sugar content of skin and of muscle in diabetic and non-diabetic persons. HARRY C. TRIMBLE AND BENJAMIN W. CAREY, JR. *J. Biol. Chem.* 90, 655-63 (1931).—In non-diabetic individuals the true (fermentable) sugar content of skin and of muscle averaged 56 and 23 mg. %, respectively, and in diabetics 144 and 51 mg. %. The elevation of blood sugar is thus accompanied by a marked absolute increase of the quantity of sugar in the skin and a much smaller elevation in muscle. In diabetics who had received insulin at short intervals previous to operation, the sugar content of the muscles averaged almost the same as in the group in which externally administered insulin was no longer believed to be exerting an effect. These results are similar to those obtained by Folin, Trimble and Newman (*C. A.* 22, 267) using experimental animals.

A. P. LOTHROP

The oxidation by potassium ferricyanide of certain constituents of the serum in anemia. G. PAYLING WRIGHT AND BARBARA ARTHUR. *J. Biol. Chem.* 90, 757-69 (1931).—Unsaturated fatty acids are oxidized by $K_3Fe(CN)_6$, and when these substances are present in the blood in any quantity as in anemia, the Haldane method for blood O_2 determinations is unsatisfactory and recourse must be made to Van Slyke's modification in which the ferricyanide-serum mixtures present a much smaller surface of contact with the O_2 , which in turn is at a low partial pressure. The progress of an anemia produced by hemorrhage is accompanied by an increase in the I absorption value of the serum fats and by an almost quantitative similar increase in the oxidizability of the serum by $K_3Fe(CN)_6$. After oxidation of the serum the I absorption values of the serum fats fall. The oleic acid present in the serum is probably only in a minor measure responsible for the oxidation reaction with $K_3Fe(CN)_6$, other more highly unsaturated fatty acids accounting for a greater part of it.

A. P. LOTHROP

The increase in production of tetanic antitoxin by addition of non-specific substances to the antigen. G. RAMON AND E. LEMETAYER. *Compt. rend. soc. biol.* 106, 71-4 (1931).—The addition to the antigen of certain non-specific substances, such as alum, $CaCl_2$ and tapioca, exerts considerable stimulation on the development of antitoxic immunity and on the production of specific antitoxin.

ALBERT L. RAWLINS

Tests on the relation between the increase of production of antitoxin and the hyperleucocytosis provoked. V. SORRIZ. *Compt. rend. soc. biol.* 106, 77-9 (1931).—The production of tetanic antitoxin is greatly increased by the addition to the antigen of a substance capable of producing at the same time a strong leucocytosis (for example, alum, $CaCl_2$, and tapioca). The antitoxin production is greatest when the leucocytosis reaches its highest level.

ALBERT L. RAWLINS

The ionic equilibrium of serum. Relations between the concentration of the salts

and the equilibrium of the albumin-globulin system. P. Lecomte du Noüy. *Compt. rend. soc. biol.* 106, 85-7(1931) —Results are given which were obtained in studying quantitatively the phenomenon of pptn. following the diln. of salt concn. by the addn. of distd. water to normal heated and unheated, serum. ALBERT L. RAWLINS

Icteric bradycardia with hyperpotassemia. DUMITRESCO-MANTE, DULUCEA and IONESCO CRAIOVA. *Compt. rend. soc. biol.* 106, 116-7(1931) —Four cases of icteric bradycardia are reported in which the blood modifications were characterized by hyperpotassemia, two of the cases being accompanied also by hypercholelmia. A. L. R.

pH and antigen of McIntosh in the Bordet-Wassermann reaction. B. ZABLOCKI and S. SIERAKOWSKI. *Compt. rend. soc. biol.* 106, 129-30(1931) —The action of the antigen of McIntosh is not affected by variations in the pH between 3.5 and 9.0. When the reaction is acid a cloudiness appears, which diminishes as the soln. is made alk. At pH 3.5 floccules appear after centrifuging, the clear, supernatant liquid does not give a reaction with the positive serums, on the other hand, the sediment gives a pronounced reaction. These floccules appearing at pH 3.5 may be divided into large and small floccules by centrifuging, the latter giving the stronger reaction with serum. ALBERT L. RAWLINS

pH and the reaction of Kahn. S. SIERAKOWSKI and B. ZABLOCKI. *Compt. rend. soc. biol.* 106, 130-1(1931) —The emulsion of the antigen of Kahn, properly mixed with a soln. of physiol. serum, has a pH of 5.8-6.5. After acidification, starting from about pH 5.0 the antigen gives a positive reaction with a negative serum (i. e., a non specific result) this is explained by the fact that the antigen itself gives a sediment which is insol. after neutralization by the serum. In this respect the antigen of Kahn differs fundamentally from the antigen of Sachs. The antigen of Kahn made alk. with NaOH gives sp. but more feeble results, thus acting like the antigen of Sachs. Acidification or alkalization of the antigen of Kahn has no influence on the final pH of the reaction, the buffering power of the antigen being quite strong but weak in comparison to the great buffering power of the serum, the mixt. of the antigen and serum causes neutralization. The best results have been obtained with emulsions of Kahn having pH 6.0-6.5. An antigen giving bad results can be corrected by the addn. of 0.5% of AcONa directly to the alc. soln. of antigen. The action of serum on the reaction of Kahn shows positive results in all cases after acidification starting at pH 5.5, after alkalization to pH 9.0 more feeble results are obtained with the positive serums, starting with pH 10.6 all positive serums give negative results. ALBERT L. RAWLINS

pH and the reaction of Sachs-Georgi. S. SIERAKOWSKI and J. MELLAK. *Compt. rend. soc. biol.* 106, 132-3(1931) —See C. A. 25, 739. A. L. RAWLINS

Physicochemical studies in gallstone formation. I. The method of the determination of the streaming potential and the applicability of Schulze-Hardy's rule. Appendix. The relation between the ζ -potential and coagulation. NONOKU FUKUTANI, TSUKASA KUROKUCHI and YOSHIO ASODA. *Japan J. Gastroenterol.* 2, 148-57 (1930) —An app. is described for measuring the $e \cdot m \cdot f$ set up by the flow of aq. solns. of electrolytes through a glass capillary under pressure. The $e \cdot m \cdot f$ is proportional to the pressure but independent of the diam. and length of the capillary, within the limits of the expts. The streaming potential decreases in the order of KCl , BaCl_2 , AlCl_3 , and $\text{Th}(\text{NO}_3)_3$ in the same mol. concn. (cf. C. A. 22, 4345). Lyophobic sols. are coagulated by a soln. of electrolyte when the streaming potential of the latter becomes sufficiently low. This corresponds to a high concn. of univalent electrolytes, a lower concn. of bivalent, and an even lower concn. of trivalent and quadrivalent electrolytes. Crystal violet behaves like the latter. II. The lyotropic series and the influence of acids and alkalis. Appendix. The relation between the ζ -potential and adsorption. *Ibid.* 158-61 —The streaming potential of cations of the same valence decreases in the order Li , Na , K , Rb , NH_4 , and Cs —and Mg , Ca and Ba , all tested as chlorides. The potential of HCl attains a much higher max. with increasing concn. than any of the salts tested. The max. potential of NaOH is much higher than that of HCl . The order for the adsorption of univalent cations (cf. Oden, C. A. 5, 3535) is the same as that for their potential, increasing from Li to Cs . III. The influence of anions and halogen group. Appendix. Physicochemical consideration of gallstone formation. *Ibid.* 162-5 —The streaming potential decreases in the order $\frac{1}{2} \text{K}_2\text{Fe}(\text{CN})_6$, $\frac{1}{2} \text{K}_2\text{Fe}(\text{CN})_6$, $\frac{1}{2} \text{K}_2\text{SO}_4$ and KCl , as well as I , Br , Cl , combined with either Na or K . An increase in the Ca content of bile may be expected to decrease the stability of the cholesterol sol. C. T. S.

Studies in carbohydrate metabolism. II. Glucose-lactic acid cycle in diabetes. HAROLD E. HINWICH, WM. H. CHAMBERS, V. D. KOSKOFF and L. H. NATHAN. *J. Biol. Chem.* 90, 417-26(1931), cf. C. A. 23, 4729, 24, 1412 —Analyses for lactic acid and glucose contents of the blood of depancreatized and of phlorizinized dogs indicate that

during exercise, and also in rest, the muscle adds lactic acid to the blood and removes hexose from it, while the liver acts reciprocally. In resting depancreatized dogs the muscle adds approx 8 mg % of lactic acid to the blood and removes 17 mg % of hexose; the liver removes 10 mg % of lactic acid and adds 23 mg % hexose. Correction for the loss in the lymph and comparison of equal vols of blood show that the liver absorbs the whole of the lactic acid added by the muscle, but the muscle in return absorbs only 13 mg % of the total 21 mg % of hexose added by the liver. Most of the remaining hexose is lost in the urine. The interdependence of muscle and liver in muscular activity and recovery, both in normal and diabetic animals, is emphasized. K. V. THIMANN.

The acid-base equilibrium of the blood in pathological conditions. I. Changes observed in the toxemias of pregnancy. EDWARD MUNTWYLER, NATALIE LIMBACH, ARTHUR H. BILL, AND VICTOR C. MILES. *J. Biol. Chem.* 90, 607-17(1931).—In 26 cases of toxemias of pregnancy a lowering of the alk. reserve of the plasma, similar to that of normal pregnancy, was observed. There was a corresponding decrease in total base from 153.8 to 144.7 millimols per l. (cf. Oard and Peters, *C. A.* 23, 2469). The bicarbonate increased considerably *post partum*, and in many cases also the total base. The lowering of alk. reserve is ascribed to hyperventilation. K. V. THIMANN.

Glucose curve in malignant tumors, before and after surgical treatment and radium therapy. G. CRAUTI AND P. BENSO. *Arch. ital. chir.* 27, 82-94(1930), *J. Am. Med. Assoc.* 96, 149.—Ten days after surgical intervention and after radium therapy there was an approach and later an actual return to the normal values of the glucose curve. R. C. WILLSON.

Calcium, magnesium, potassium, sodium and chlorine in blood after cranial injuries. G. SICILIANI. *Arch. ital. chir.* 27, 95(1930), *J. Am. Med. Assoc.* 96, 149.—Cranial injuries more or less grave (commotion, fracture of the base) cause a const. but transitory disturbance, for not more than 36 hrs., of the electrolytic equil., characterized by a diminution in Ca and an increase in K and P, while the remaining electrolytes behave in an inconst. manner. These changes were more pronounced when the trauma affected the area of the base of the cranium. R. C. WILLSON.

H—PHARMACOLOGY

A. N. RICHARDS

Use of 7% carbon dioxide and 93% oxygen in the treatment of carbon monoxide poisoning. C. K. DRINKER AND J. SNAUGHNESSY. *J. Ind. Hyg.* 11, 301-14(1929).—The treatment of persons suffering from acute CO poisoning is best carried out by the inhalation during the first 5 to 20 min. of mixts. contg. 7% of CO₂ and 93% of O₂, followed by the usual mixt. contg. 5% of CO₂ and 95% of O₂. B. C. A.

Solution of dextrose and sodium chloride for obliterating varicose veins. HOWARD M. KERN. *Ann. Surgery* 93, 697-706(1931).—A study of 464 cases in which an equal mixt. of 50% dextrose and 30% NaCl solns. was used almost exclusively. R. B.

The action of gold preparations in tuberculosis. FRANZ WIRL. *Z. Tuberk.* 65, 489-91(1930).—The author used Aurophos and Lopion in two cases of skin tubercules originating hematogenously, and there resulted new lesions from this Au treatment. It is, therefore, concluded that the Au preps. are not parasitotropic, and it is suggested that the chemical industries seek new combinations to overcome this deficiency. H. J. CORPER.

An attempt to use leucocytes as a mechanical vehicle for transporting therapeutic agents to the lungs. PHILIPP SPANIER. *Beitr. klin. Tuberk.* 73, 210-31(1929).—S. verified the observations of Christeller and Eisner that heterogenous living leucocytes introduced into the blood stream of an animal mostly are retained in the pulmonary capillary while they pass the capillaries of all the other organs. This idea is utilized for transporting therapeutic agents. The leucocytes were obtained from dogs in which sterile turpentine oil abscesses were produced. The substances to be phagocytized were introduced either *in vivo* by injecting into a 3-day old abscess and removing the pus one quarter hr. after introduction or *in vitro* by shaking the leucocytes at 37° with an equal vol. of isotonic soln. prepd. in 0.85% NaCl and 1.5% Na citrate used in equal amts. The suspension was given in 1 cc. amts. to rabbits or dogs intravenously. In these expts. trypan blue, iron preps., I and contrast media such as BaSO₄ were utilized. H. J. CORPER.

Benzene-oil treatment in tuberculosis. JONAS KAIRIUKSTIS. *Beitr. klin. Tuberk.* 73, 456-9(1930).—The author used a 30% sterile soln. of benzene in olive oil by intramuscular injection in amts. of 0.1 to 0.2 cc. This caused a focal reaction in tuberculous individuals which resulted in a favorable action upon the course of the disease. The

results in surgical tuberculosis appeared to be better than those in pulmonary tuberculosis. The clinical data are given in summarized form. H. J. CORPER

Therapeutic experiments with ascarirole. JOHANNES ZSCHOCKER. *Arch. Schiffs-Tropen Hyg.* 34, 472-80(1930) — Good results are reported in the treatment of *Necator* and *Ascaris* infections. H. EAGLE

The treatment of children's dysentery with Rivanol. D. TRITTEL. *Arch. Schiffs-Tropen Hyg.* 34, 481-6(1930) — Rivanol taken in the mouth in doses of 5 mg 3-6 times daily has a marked therapeutic effect upon amebic dysentery. If unsuccessful after 5-7 days, it should be supplemented by 2-3 injections of emetin. H. EAGLE

The treatment of surra with benzenearsonic acid "Höchst 4002." OTTO NIRSCHL AND I. K. HAWO-ROENTGE. *Arch. Schiffs-Tropen Hyg.* 34, 662-9(1930) — Unsatisfactory results were obtained in the treatment of experimentally inoculated rats, mice and guinea pigs. H. EAGLE

The chemotherapeutic effect of gold compounds upon mice inoculated with *Spirochaeta crociduræ*, with particular reference to treatment with mixtures of neoarsphenamine and Solganal. W. MEYER. *Arch. Schiffs-Tropen-Hyg.* 35, 97-107(1931). — Solganal is much more effective than neoarsphenamine as a sterilizing agent in inoculated mice. M. recommends the use of a mixt. of the 2 in the treatment of recurrent fever and suggests a possible application in the treatment of syphilis. H. E.

Cholosulin. L. R. GORM. *Arch. Verdauungskrankh.* 43, 1-24(1930) — Cholosulin (an addn. compd. of insulin with the Na salt of desoxythioic acid) is supposed to be an insulin which can be taken by mouth. G. found it effective in only 2 out of 16 cases of diabetes. H. EAGLE

Can the retention of foods in the human stomach be affected by the administration of various kinds of beer or of their carbon dioxide and bitter contents? A. BICKEL AND F. FLEISCHER. *Arch. Verdauungskrankh.* 43, 28-42(1930) — Beers cause a delayed emptying time of the stomach because of their alc. content, rather than their CO₂ or bitters contents. H. EAGLE

The effect of ephedrine upon gastro-intestinal activity. F. TRAKHTENBERG. *Arch. Verdauungskrankh.* 43, 69-95(1930) — Ephedrine inhibits the motility of the stomach and intestines for 3-4 hrs., preceded by a short period of excitation (8-10 min.). The gastric acidity is unaffected, but the total Cl⁻ is increased. Its therapeutic use is indicated in gastric or duodenal ulcer, gastric hypertonicity and hypermotility, spastic obstruction, etc. H. EAGLE

The insulin-sensitivity of nondiabetic humans and its relationship to the vegetative nervous system. EADSR STRAMM. *Arch. Verdauungskrankh.* 43, 104-24(1930). — The intravenous injection of 1/10 unit of insulin into normal subjects causes a maximal fall in blood sugar of 20.5 mg %. In vagotonic subjects, this may be as high as 43.6 mg %; in hypothyroidism it averages 47.2 mg %. On the other hand, in subjects with increased sympathetic tone (e.g., with hyperthyroidism) there is less sensitivity to insulin with an av. fall of only 16.9 mg % under the same conditions. H. EAGLE

The ingestion of vinegar and lemon juice in their relationship to the morphological character of the blood. ADOLF BICKEL AND GERHARD LISS. *Arch. Verdauungskrankh.* 43, 158-66(1930), cf. C. A. 24, 2169 — Continued ingestion has no harmful effect upon the blood picture. H. EAGLE

Excretion of alcohol in gastric juice following the rectal administration of alcohol. ARTHUR LUKAS. *Arch. Verdauungskrankh.* 43, 332-45(1930). H. EAGLE

Histamine. ANTON KERRER. *Klin. Wochschr.* 9, 2137-40(1930), cf. C. A. 24, 5378 — The possible role of histamine in intestinal intoxication is discussed. As little as 1/2 mg. injected subcutaneously into humans causes an increased secretion of gastric juice contg. both acid and pepsin. Intravenous injection is effective only if it is given very slowly. In approx. 1/2 the cases which are diagnosed as achylia gastrica, histamine causes an active gastric secretion. It is to be emphasized that a pos. histamine test does not exclude organic gastric lesions, notably cancer, nor does an absence of secretion mean irreparable anatomical damage. This stimulating action of histamine upon gastric secretion is due to a direct action on the secreting cells. Histamine is found in all tissues, and in very high concn. in the lungs, presumably as an intermediary product of metabolism. The close similarity between anaphylactic shock and the pharmacological effects of histamine suggests the etiological role of the latter. H. EAGLE

Chemistry of uroselectan. A. BIRN AND C. RATH. *Klin. Wochschr.* 9, 2297-8(1930) — 2 Pyridone-5-arsonic acid was found to be less toxic than any known org. linkage of As, in particular, the arsonic acids of the benzene ring. I introduced instead of the arsonic acid also resulted in a compd. of very slight toxicity, its Na salt was intro-

duced in 1927 as a therapeutic agent ("Selectan") in the treatment of streptococcal mastitis in cows. The Na salt of 1,2-dihydro-5-iodo-2-keto-1 pyridineacetic acid is uroselectan, sol up to 35%, the lethal dose of which is 3.37 g. per kg when injected intravenously into the rat. H. EAGLE

Cholesterol and lecithin in the water and acid-base economy. RUDOLF DEGWITZ. *Klin. Wochschr.* 9, 2336-44(1930).—Cholesterol and lecithin have antagonistic and mutually neutralizing effects upon the hemolysis of red cells by water, acid or alkali. In general, cholesterol increases the resistance of the cell, possibly because, being a hydrophobic substance, it promotes a protoplasmic change in the nature of a water-in-oil dispersion, in which effect it would be, and is, aided by Ca^{++} and Mg^{++} . Lecithin, on the other hand, being hydrophilic, might promote an oil in water dispersion, which would be aided by Na^+ and K^+ . Cholesterol injected into rabbits causes a transport of H_2O and Cl^- from the tissues into the blood and urine, with polyuria, and very high blood and urine $[\text{Cl}]$. Hyperlecithinemia, on the other hand, causes a low plasma and urine $[\text{Cl}]$, and a retention of H_2O and Cl^- in the tissues. Paradoxically, the prolonged peroral administration of cholesterol causes water and Cl^- retention, and decreased urinary excretion. H. EAGLE

The effect on the experimental animal of injection of a liver extract, prepared according to the method of Gänselen. W. EDERLE AND H. KRZICH. *Klin. Wochschr.* 10, 25-6(1931).—The ext. had a very favorable effect in white rats with a *Bartorella* anemia, whose spleen had been removed. H. EAGLE

Relationships between insulin and adrenaline in the human organism. BERNHARD KUCLEMANN. *Klin. Wochschr.* 10, 59-62(1931).—If the blood sugar falls below 60 mg % following the intravenous injection of insulin, there is a considerable rise in systolic pressure and fall in diastolic pressure, the pulse pressure being doubled or even tripled. As the blood sugar returns to normal, there is a concomitant return of blood pressure to normal values. The administration of glucose prevents the blood pressure change, which is only observed if the blood sugar falls below 60 mg %. Simultaneous with the rise in pressure there is a moderate leucocytosis, with a particular rise in lymphocytes; and, in patients with an enlarged spleen, there is a marked decrease in its size. Since all three of these phenomena (blood pressure, blood count and splenic changes) are observed following the injection of adrenaline, K. concludes that in insulin hypoglycemia there is a mobilization of adrenaline as a compensatory factor, tending to restore the blood glucose to its normal values. H. EAGLE

Adenosinephosphoric acid and cardiac activity. A contribution to the component of muscle extract which affects the heart. HANS ROTHEMANN. *Klin. Wochschr.* 10, 67(1931).—Adenosinephosphoric acid causes slowing of the heart rate, and induces a more regular and stronger action in pathological conditions. Since it can be demonstrated in muscle exts., it is concluded that it is the active constituent in such exts. as regards cardiac function. H. EAGLE

The analysis of the action current curve of the heart into chemical phases. I. Relation between the action current and the formation of lactic acid. RUDOLF SIEGEL AND Z. UNNA. *Klin. Wochschr.* 10, 72-3(1931).—Monooxoacetic acid is known to prevent the formation of lactic acid in striated muscle. In mice and rabbits small doses cause a change and broadening of the T-wave of the electrocardiogram while larger doses cause its complete disappearance without affecting the Q R S complex. S. and U. suggest that the T-wave and the formation of lactic acid are due to the same process; moreover, that the formation of lactic acid begins 0.05 sec. after the beginning of contraction (Q R S complex) and continues through systole. II. Contractions without after-waves in the electrocardiogram. *Ibid.* 171.—After poisoning with $\text{CH}_3\text{CO}_2\text{H}$, the ventricles continue to contract despite the absence of the T-wave in the electrocardiogram. H. EAGLE

The excretion curve of iodine from the blood following the administration of sodium tetraiodophenolphthalein under normal and pathological conditions in the animal experiment. H. EITEL AND A. LOESER. *Klin. Wochschr.* 10, 109-11(1931). H. E

The locally limited action of hormones, particularly that of pituitrin, upon the water exchange. ARISTO KISS. *Klin. Wochschr.* 10, 152-5(1931).—Pituitrin has an effect upon the connective tissue of the skin, causing an increased water absorption. The degree of its effect depends upon the conditions of the skin, which is apparently under the control of the liver. H. EAGLE

Comparative actions of sympathomimetic compounds: Influence of cocaine and certain related compounds upon the actions of a group of sympathomimetic amines. M. L. TAINTER. *Quart. J. Pharm. Pharmacol.* 3, 354-68(1930); cf. C. A. 24, 894.—The cocaine sensitization-desensitization was applied to a group of sympathomimetic

amines Dihydroxyphenylethylamine and dihydroxyphenylpropylamine were sensitized by cocaine and reversed by ergotoxine, and thus were sympathicotropic, in common with the other catechols hitherto studied. The substitution of nitro, chloro or carboethoxy groups in the para position of phenylethylamine or the replacement of the Ph by a thiophene group produced precursor compounds which were desensitized by cocaine, and hence were pseudo-sympathicotropic. *p*-Carboxyphenylethylamine, Phyl IN and "salvamin," a trihydroxy product, were inactive on blood pressure. Paeucaine, benzoyl-pseudotropine (tropacocaine) and β -eucaine, which are closely related to cocaine chemically, did not reproduce the cocaine sensitization-desensitization. Cyanides, by interfering with the oxidative processes, depressed the response to adrenaline, tyramine and pituitary equally, and therefore did not resemble the cocaine action. Also, smooth muscle depression sufficient to lower the blood pressure, from large doses of papaverine, did not depress the action of these precursor drugs. It therefore seems unlikely that interference with oxidative processes or smooth muscle depression can be essentially involved in the phenomena. Cocaine increased the hyperglucemic responses of rabbits to adrenaline, and decreased those to ephedrine. The typical cocaine sensitization-desensitization phenomena can, therefore, be seen in functions not involving a smooth muscle response. W O E

An unsuccessful abortion attempt or suicide by means of arsenic. A SARTORI. *Chem Ztg* 55, 105-6(1931).—Investigation into the death of a pregnant woman at first presented a picture of ergot poisoning. Later a bag containing a gray powder was found. Analysis showed powder C and As_2O_3 . RUSSELL C. EMM

Effect of various concentrations of nicotine in tobacco on the growth and development of fowls. J E HUNTER AND D E HALEY. *Penna Agr. Expt. Sta., Bull* 258, 9(1931). Abstract 434 *Annual Report of the Director*.—When a high nicotine strain of tobacco is used in the feed of chicks, remarkably large amounts of nicotine can be tolerated. By building up a tolerance to tobacco in baby chicks, as much as 1.2% of pulverized tobacco can be included in the ration without noticeable injury. In fact there appears to be a stimulation due to the tobacco and better growth gains are made than by the controls. It is believed that continuous tobacco treatment will effectively control worm infestation. C R. FELLERS

Pharmacological action of irradiated proteins and protein derivatives. P. HOLTZ. *Arch. ges. Physiol. (Pflügers)* 226, 57-77(1931).—Serum irradiated with ultra violet light has a dilator action on the blood vessels of the frog. Irradiated euglobulin, glycine, leucine, histidine and tyrosine show the same action. The action of the irradiated serum is attributed to HCNO, AcNH and imidazole groups, whose presence in such serum was demonstrated. The stimulating effect of irradiated sera on the intestine was similarly attributed to the presence of histamine or another imidazole derivative. A G

The effect of irradiated serum on sugar elimination by the isolated liver of warm-blooded animals. P. HOLTZ. *Arch. ges. Physiol. (Pflügers)* 226, 578-84(1931).—Irradiated serum markedly affects sugar elimination by the liver of guinea pigs, cats and dogs. In most cases the elimination is reduced. ARTHUR GROLLMAN

The changes in the stimulation time-tension curves by narcosis, cold and veratrine. CHUNG LING HOU. *Arch. ges. Physiol. (Pflügers)* 226, 676-83(1931). A G

Mode of action of compounds of insulin with bile acids on subcutaneous administration. R. STEPHAN. *Deut. med. Wochenschr.* 56, 83-81(1930).—Sodium desoxycholate, coupled with insulin, shows the same action when administered subcutaneously as after oral administration. The clinical action of the combination is, however, different from that of insulin. ARTHUR GROLLMAN

The present status of chemotherapy, particularly in the treatment of tropical diseases. C. SCHILLING. *Deut. med. Wochenschr.* 56, 955-8(1930).—An address dealing with the modern trend of chemotherapeutic research. ARTHUR GROLLMAN

Cholosulin and the daily blood sugar curve. RICHARD KERN. *Deut. med. Wochenschr.* 56, 1341-2(1930).—Cholosulin, a combination of insulin and the Na salt of desoxycholic acid, produced no change in the blood sugar curves of diabetics. A G

The initial insulin hyperglucemia. A. V. KORJAK. *Deut. med. Wochenschr.* 56, 1633(1930).—The initial hyperglucemia which follows the injection of insulin is attributed to a demobilization of sugar in response to the rapid depletion of the sugar in the muscles. ARTHUR GROLLMAN

The effect of saponins on the basal metabolism in Basedow's disease. A. BORNSTERN. *Deut. med. Wochenschr.* 56, 1661-2(1930).—The increased metabolism in thyroid toxicosis is due to nervous stimulation and to an increased metabolic activity on the part of the cells. The former fraction of the metabolic increase is abolished by the use of saponins, the latter remains unaffected. ARTHUR GROLLMAN

Is lactic acid absorbed through the skin? E. E. DAUKE. *Deut. med. Wochschr.* 56, 1869-70(1930).—No evidence is available to indicate that lactic acid or other nutrients may be absorbed through the skin.

ARTHUR GROLLMAN

Distribution between blood plasma and red blood corpuscles of some substances used in therapeutics. R. FADRE. *Bull. soc. chim. biol.* 12, 954-64(1930), cf. *C. A.* 25, 1591.—In dog blood, diethylmalonylurea and allylpropylmalonylurea showed a distribution between red corpuscles and plasma of 6 mg 3.6 mg and 19.1 mg 19.3 mg, resp., per 100 g blood. Quinine and hydrastine were about equally distributed between corpuscles and plasma but were retained longer by corpuscles. C. G. KING

Pharmacological influence of the substitution of a methyl group in the β -carbon of methylaminomethyl(p -hydroxyphenyl)carbinol. RAYMOND-HAMET. *Compt. rend.* 192, 450-2(1931).—The action of methylaminomethyl(p -hydroxyphenyl)carbinol has been studied in double suprarenalctomized dogs which have been treated with cocaine or yohimbine. Its action has been compared with that of p -hydroxyephedrine and ephedrine. The substitution of a methyl group in the β position has endowed the drug with an adrenaline like action, which it lacked. G. H. W. LUCAS

Investigation on the pharmacological action of the vessels of the pigeon's wing. 1. The influence of the changes in concentration of sodium, calcium and potassium chlorides, as well as in the hydrogen-ion concentration in the perfusing Ringer solution, TOMONORI MANABE. *Folia Pharmacol. Japon.* 11, 225-37, Breviaria 17(1931).—The vessels in the pigeon's wing were perfused with Ringer soln. in which the chloride content and pH were varied. Excess NaCl increases persistently the discharged quantity but a large excess causes a decrease. The vessels are dilated and the osmotic pressure is increased. The decrease in flow is due to a swelling of the endothelium in the small vessels, especially the capillaries. The tissues also are affected. The addn. of $CaCl_2$ causes a dilation of the vessels with small amts and a constriction with larger quantities. Absence of $CaCl_2$ in Ringer soln. produces a dilation. The addn. of KCl produces a dilation or a constriction, depending on the quantity. Alk. Ringer soln. causes a dilation, as does a highly acid one. G. H. W. LUCAS

The experimental investigation of Panax ginseng. IV. The virulence of some nerve poisons and the poison phenomena brought about in rats fed on Panax ginseng. PIUNGKI MIN. *Folia Pharmacol. Japon.* 11, 238-55, Breviaria 18(1931), cf. *C. A.* 24, 5869.—The nerve poisons adrenaline, physostigmine and pilocarpine act synergistically, while atropine and adrenaline act antagonistically in ginseng fed rats. V. The comparative study on the chemical constituents and the general action of Panax ginseng and Panax quinquefolius. *Ibid.* 256-60, Breviaria 19(1931).—The chief constituents of each plant were quantitatively prepd. Panax ginseng contains more panacene, panax acid, phytosterol and glucoside than Panax quinquefolius. ALC. exts. of both plants injected into mice produced tail symptoms (the tail being carried erect), convulsions and death. Panax ginseng exts. were more potent. G. H. W. LUCAS

The antiphlogistic action of drugs on inflammation produced by streptococci. IRIKOSHI SUZUKI. *Folia Pharmacol. Japon.* 11, 276-85, 287-98, Breviaria 20-1(1931); cf. *C. A.* 25, 740.—Antipyrine, quinine, erycon, salicylic acid, sinomenine, casein, rabbit serum, streptococcus serum and streptococcus vaccine exert an antiphlogistic action on the inflammation produced by streptococci-salt soln. in the rabbit eye. The action of antipyrine, quinine, erycon and sinomenine on inflammation produced by streptococci is weak, while that of salicylic acid, casein, rabbit serum and staphylococcus is scarcely antiphlogistic. The action of the drugs is stronger on the inflammation produced by streptococci than on that produced by chem. or phys. stimulation. Drugs act but slightly on the inflammation produced by cold in the rabbit eye. G. H. W. LUCAS

The pharmacology of benzoylglucuronic acid. SHINZO SAKATA. *Folia Pharmacol. Japon.* 11, 299-324, Breviaria 21-2(1931).—Warm blooded animals are protected from poisoning by chloral hydrate, camphor, phenol, naphthol, borneol, antipyrine and morphine by the formation of a conjugated compd. with the glucuronic acid in the body. The conjugated glucuronic acid in the urine is thus increased. The conjugated acid belongs to the glucoside type, whereas the benzoyl derivs. are of the ester type. Phenylglucuronic acid, urochloralic acid and the conjugated ester type await investigation. The action of benzoylglucuronic acid was measured on rabbits and toads. On the hearts of these, tonus is increased and contraction height decreased by small doses, large doses cause a fall in tonus. The rate is unchanged and the heart stops in diastole. The action appears on the muscle itself. In small doses the drug causes a dilation of vessels in the rabbit ear and in the frog extremities. In large doses a constriction follows. It has some diuretic action in rabbits and toads, kidney vessels are dilated. A transient bydremia follows its administration. The diuretic action

is present in the toad after the ligation of the kidney portal vessels. Small doses stimulate and toxic doses paralyze isolated rabbit organs. G. H. W. Lucas

A contribution on the pharmacology of phenylpyrrolic acid. TADASHI IASUI AND HIROTO MATSUHIMA. *Folia Pharmacol Japon* 11, 261-75; Breviaria 20(1931).—Various amino acids and their intermediate metabolic products possess some influence on certain hormone actions in animal organs. The intermediate product of phenylalanine, phenylpyrrolic acid, was studied. On perfused toad livers, it increased sugar. It produced no action on the pupil. On the surviving rabbit intestine it had a depressing action. On subcutaneous injection into rabbits, it increased blood sugar and was not synergistic with adrenaline. On the isolated hearts of toads and rabbits it had a paralyzing effect, its action being on the muscle itself. G. H. W. Lucas

The relationship between the action of some cardiac tonics on the embryonic chicken heart and its stage of development. SAKURO KIRIARA. *Folia Pharmacol Japon* 11, 25-31; Breviaria 22, 3(1931).—The action of coramine, digitalein, cardiazole and aconitine on 2 to 10 day old chicken embryo heart has been followed. A 2 day old embryo is stimulated by these heart tonics. The sensitivity increases up to max where it remains unchanged for about 10 days. The action of these drugs on the stage of development was also studied. G. H. W. Lucas

The influence of various conditions on the resorption of sodium salicylate from the pleural cavity as well as on the manner of resorption of these drugs. II. MASAYOSHI OGAWA. *Folia Pharmacol Japon* 11, 334-45; Breviaria 23(1931), cf C. A. 25, 1286.—The right lung was inflated with AgNO_3 . 0.1 g. Na salicylate per kg. was injected in the pleural cavity and the rate of resorption followed. Relatively slight inflammation resulted in good absorption, but marked inflammation slowed up the rate on account of damage of the capillaries. Respiratory movements facilitated resorption. Secretion of exudate following damage to vessels lowers the concn. of the applied material. III. *Ibid* 345-58; Breviaria 24-5.—The lungs were treated in the normal and in an inflamed condition, and the blood in the lung veins was examined for Na salicylate. Low-grade inflammation arrests resorption, whereas high grade hinders it. A closed pneumothorax hinders resorption. The disturbance in resorption is more marked in the pleura pulmonalis than in the pleura parietalis. G. H. W. Lucas

The relationship between the vomiting, blood-sugar- and urea-acid-regulating centers from a pharmacological standpoint. II. The influence of emetics on the urea acid excretion. MORIYU MIYU. *Folia Pharmacol Japon* 11, 259-69; Breviaria 25(1931), cf C. A. 25, 1284.—Comparative expts were performed by injecting dogs subcutaneously with emetics and following the urea acid excretion. Morphine-HCl was used as a control. Following the injection of apomorphine-HCl, emetine HCl, wine of Sb, CuSO_4 , and eryoin the urea acid excretion was increased over the control. Na salicylate and antipyrine lowered the temp. and blood sugar stimulated vomiting and promoted the formation of urea acid. The centers regulating vomiting, blood sugar and urea acid are in close relationship with each other. G. H. W. Lucas

Effects of insulin and adrenaline on the blood acetone in fasting children. LARRY SALOMONSON. *Am J Diseases Children* 40, 718-24(1931).—The antiketogenic action of insulin appears to be secondary to the influence of insulin on sugar metabolism. The administration of insulin to children whose blood sugar has been reduced by fasting reduces the concn. of acetone in the blood, the reduction being less marked the lower the sugar concn. The administration of adrenaline increases the concn. of acetone, the increase being more marked the lower the concn. of blood sugar. The injection of adrenaline is usually followed by an increase in the fatty acid content of the blood. E. R. MANN

The effect of phosphorus in rickets. I. Röntgenologic changes in rickets following administration of phosphorus. EDWARD L. COOPER. *Am J Diseases Children* 40, 941-67(1933).—The administration of a combination of small quantities of elementary P and cod-liver oil to infants with severe cases of rickets has a much greater therapeutic effect than cod liver oil alone. The administration of P alone is not sufficient to cause the healing of severe cases of rickets. The studies were made during the late fall and winter months. II. The chemical changes in the blood in rickets following administration of phosphorus. *Ibid* 1177-92.—The administration of a combination of cod liver oil and P is followed by a marked increase in the Ca and P contents of the blood serum and in the product of the two. The administration of cod liver oil alone to infants with severe rickets is followed by a slow but definite increase in the product of the Ca and P contents. The administration of P alone causes no significant changes in the Ca or P contents or in the product of the two. The healing of rickets does not appear to be associated with changes in the pH of the blood serum. E. R. M.

Coffee as an intestinal disinfectant. ERNST G. DRESSEL AND HAROLD LOTTE. *Arch Hyg* 104, 144-55(1930) —Broth infusions of finely ground coffee or preps. from which the caffeine has been removed do not permit the growth of typhus bacilli or cholera vibrio, they inhibit the growth of Flexner bacilli and kill Shiga-Krusse bacilli. Finely ground coffee appears to be a more effective agent in its action against the pathogenic microorganisms of the intestinal tract than the com. preps. of C.

E. R. MAIN

The blood in experimental benzene poisoning. L. P. BRYULLOVA, A. S. BRUSILOVSKAYA, N. V. LAZAREV, M. P. LYUBIMOVA AND D. I. STALSKAYA. *Arch Hyg* 104, 225-38(1930) —No marked or characteristic chem. or histologic changes appear in the blood of lab. animals poisoned with benzene vapor.

E. R. MAIN

Paths of absorption and excretion of sodium tetraiodophenolphthalein. A. J. DELARNO. *J. Lab. Clin. Med.* 16, 329-40(1931) —Na tetraiodophenolphthalein appears to be absorbed by both the small and large intestines following oral or rectal administration to dogs. It leaves the blood within 2 to 4 hrs. after intravenous administration. The liver excretes from 60 to 70% of the dye, apparently unchanged, the large intestine about 25%. From 5 to 10% is eliminated in the urine partly in an more state. Some of the dye may be rendered insol. by contact with pancreatic juice or fat present in the small intestine. The results are based upon the detn. of I after oxidation of the blood, urine, etc.

E. R. MAIN

The effect of calcium-precipitating substances on the ciliated epithelium of the maxillary sinus. W. F. WYVER. *J. Lab. Clin. Med.* 16, 341-6(1931) —The movement of the cilia of the ciliated fragments of ciliated epithelium of the maxillary sinuses of rats and rabbits ceases when aluminum, citrate or oxalate solns. are added to the medium in which the tissues are immersed.

E. R. MAIN

Some effects of amytal anesthesia. J. M. OLIMSTED AND GEORGE GERACOSSENTI. *J. Lab. Clin. Med.* 16, 354-61(1931) & C. A. 25, 149 —The intravenous injection of amytal into dogs is not followed by an increase in the blood sugar content; of dogs that have been fed a diet of lean meat. A slight increase may follow if the dogs have been fed a diet rich in carbohydrates. The hyperglycemia which accompanies the administration of morphine is prevented when amytal is injected simultaneously with the morphine. The increase in blood sugar that normally follows a short period of asphyxia is prevented by amytal. In man, the non protein N and urea contents of the blood are markedly increased, the blood pressure is reduced and the heart rate increased after the administration of amytal. The gastric secretion of dogs is inhibited during amytal anesthesia.

E. R. MAIN

The glucose reaction with minimum quantities of adrenaline administered intravenously. L. CARRARO. *Boll. soc. ital. biol. sper.* 5, 810-3(1930) —Small variable doses of adrenaline were injected intravenously into normal persons, and into patients afflicted with various diseases. The results showed that the glucose effect of insulin is variable and is connected with the constitutional state of the subject and the sympathetic-parasympathetic equil. In some individuals there was a fall in blood sugar 5 min. after the injection and then a rise. Adrenaline, when injected intravenously, may cause appreciable changes in the glucose titer in doses as low as 1/400 mg. Patients with hepatic lesions showed no noticeable difference to the action of adrenaline in contrast to normal or to patients affected with other diseases. PETER MASTUCCI

Qualitative and quantitative studies on the penetration of the lungs by atomized drugs. POMERO SPORO. *Boll. soc. ital. biol. sper.* 5, 815-7(1930). —A method is described for studying the penetration of atomized drugs in the lungs of the dog. The substances tested were methylene blue, rosin oil colored with Sudan III, and colloidal Ag. From the results obtained, it is shown that the penetration of atomized liquids does occur if the mouth is kept open and the nares are closed; the atomized liquids may reach the alveoli of the lungs but only in small amts. and so unequally distributed that parts of the lung are not reached.

PETER MASTUCCI

Quantitative studies on the fixing power of the hepatic reticulo-endothelium apparatus. G. ORESTANO. *Boll. soc. ital. biol. sper.* 5, 818-9(1930) —Colloidal AgS was injected intravenously into rabbits. After different intervals of time the animals were killed and the Ag fixed by the liver was detd. After 9 min. 1/4 of the Ag injected was found in the liver. With longer intervals after injection, the velocity of fixation diminished, 1/4 was fixed in 2 hrs. and about 1/4 after 8 to 10 hrs. FbS gave similar results after 2 hrs. 1/4 of the Pb was found in the liver.

PETER MASTUCCI

The relation between the action on the hemopoietic tissue and fixation of colloidal copper sulfide in the reticulo-endothelial apparatus. GERLANDO LO CASCO. *Boll. soc. ital. biol. sper.* 5, 819-22(1930) —The injection of 10 cc. of an 0.5% soln. of col-

loidal CuS into rabbits resulted in the presence of 23,000 nucleated erythrocytes in 24 hrs. If however, the injection of CuS is preceded by the injection of HgS (0.3 to 1 cc of 1% soln) the erythroblastic reaction diminished in proportion to the HgS injected. India ink and lithocarmine gave similar results but HgS was the substance which produced the most complete, permanent and innocuous block of the reticulo-endothelial app. of the bone marrow, spleen and liver. **PETER MASLUCCI**

The action of quinine on the reticulo-endothelial system. **A. LEVI AND F. BELLONI** *Boll. soc. ital. biol. sper.* 5, 857-8 (1930), cf. *C. A.* 24, 2176.—Quinine in relatively high doses (1-2 mg) inhibits the development of the reticulo-endothelial system in mice, in smaller doses, (0.5 mg) equal to those used in inhibiting the development of tumors, it stimulates the growth of this system. **PETER MASLUCCI**

The resistance of the corporales after using antipyretics. **A. LEVI AND R. BASSI** *Boll. soc. ital. biol. sper.* 5, 859-60 (1930).—The repeated administration to rabbits of antipyretics (antipyrine, pyramidone, melubrin, cryogenin, maretin) in therapeutic doses (0.5 g) does not give rise to variations in the resistance of the corporales and does not cause hemolysis. Toxic or mortal doses do exert a slight change in resistance. The above antipyretics may therefore, be satisfactorily used in therapy. **P. M.**

The neuromuscular junction in curarized frogs and in frogs treated with quinine. **G. M. PICCINI AND C. DELPINO** *Boll. soc. ital. biol. sper.* 5, 861-2 (1930).—The gastrocnemius of the normal, of the curarized and of the quinine treated frog when examined microscopically showed no sp. changes in the terminal motor app. **P. M.**

The pharmacodynamic posology of quinine for frogs and the course of the aponeurotic and reflex motility. **G. M. PICCINI AND C. DELPINO** *Boll. soc. ital. biol. sper.* 5, 862-6 (1930).—Quinine HCl 2H₂O was injected into the dorsal lymph sac of (1) hibernating and (2) summer *Rana esculenta*. Doses of 0.10, 0.08, 0.05, 0.03 g. in weak solns (1/30, 1/50, 1/100) injected into frogs of 40, 30, 20, 12 g. were all lethal. The m. l. d. was 0.25 mg per g. of frog. All the lethal doses produced a depression followed by a general paralysis. Under these conditions weak elec. stimuli applied to the skin contracted the muscles energetically. The reflex motility gradually decreased but lasted up to 24 hrs. Doses approaching the m. l. d. (0.2 to 0.07 mg) produced the same toxic symptoms which were often preceded by a transitory excitation. The frogs were only apparently dead but within 24 hrs. became normal again. Moderate doses (0.005-0.04 mg) always produced excitation phenomena of varying duration and intensity. Sub-specific doses of 0.001-0.04 mg showed nothing definite but did increase muscular work. The effects of injecting concd. solns. were also investigated and are discussed. **PETER MASLUCCI**

The action of insulin on the muscle contraction of the frog. **G. BUCCARDI AND A. DE NIEDERHÄUSER** *Boll. soc. ital. biol. sper.* 5, 870-2 (1930).—Insulin injected into the frog causes an increase in the amplitude of muscle contractions and muscular work. The increase in the capacity for work coincides with a diminution of muscular glycogen. Insulin convulsions cause a diminution of muscle glycogen and a decrease in muscular work. **PETER MASLUCCI**

Experimental studies on the hypoglycemic action of beer yeast administered enterically. **AURELIO COSTA** *Boll. soc. ital. biol. sper.* 5, 878-81 (1930).—Beer yeast administered in 5-g. doses by stomach tube to rabbits, as a general rule, produced slight hyperglycemia. It did retard, however, the hyperglycemic peak usually obtained when glucose is injected intravenously into rabbits. **PETER MASLUCCI**

The action of insulin on thermoregulation. **GIULIO CERUTI AND ENRICO SAGEGNO** *Boll. soc. ital. biol. sper.* 5, 905-6 (1930).—Under the action of insulin, if the dose has been insufficient or due to the peculiar condition of resistance of the individual animal the hypoglycemia does not reach below a certain level, the internal temp. does not vary markedly. However, if the dose has been sufficiently large to produce a marked lowering of blood sugar approaching the convulsive dose (0.00-0.05%) then there is a rapid fall of temp. which is related to the glucemic titer. The fall of temp. is arrested by the injection of glucose. The temp. was detd. by a thermocouple method. **PETER MASLUCCI**

The mechanism of action of colloidal sulfur. **G. RASTELLI AND A. CASAZZA** *Boll. soc. ital. biol. sper.* 5, 919-42 (1930).—Egg albumin and blood serum have the power to transform S into H₂S, the greater the dispersion of the S, the less is the reduction to H₂S. Colloidal S when injected into the circulation exerts a double action. The portion of the S which exists in the true colloidal state det. the anaphylactic type of shock, while the portion which exists in a state of flocculation and which is present in all preparations, becomes reduced to H₂S which det. the toxic effect. **P. M.**

Radium emanations and the physiologic action of certain drugs. **P. MASCHERPA**

AND G. BELLUZZI *Boll soc ital biol* *ser* 5, 944-6(1930)—The expts. were made on frogs and mice kept in a closed container into which circulated air that had been bubbled through a soln of Ralir. Morphine-HCl was injected into frogs in 0.01 g. doses and into mice in 0.001 g. doses and strychnine sulfate into frogs in 0.00002 g. doses and into mice in 0.00005 g. doses. It was observed that in the animals treated with the Ra emanations the toxic picture brought on by the 2 alkaloids reached greater degrees of violence and appeared much more rapidly than in the control animals. Often the doses administered were lethal to the radioactivated animals. PETER MASUCCI

Biochemistry of magnesium narcosis. EMILIO MARTINI *Boll soc ital biol. ser* 5, 946-8(1930)—The narcotic efficiency of Mg salts depends on the alk. reserve of the blood. The higher the alk. reserve, the less is the Mg content, or the greater is the Ca content of blood of the awakened animal. M. makes the deduction that an increase or a diminution of the alk. reserve corresponds to an increase or decrease, resp., of the narcotic efficiency of Mg. PETER MASUCCI

The action of bulbo-capnine, injected intravenously on the local reflexes of posture of man. U. DE GIACOMO AND A. SEVERINO *Boll soc ital biol ser* 5, 965-6(1930), *cf C A* 25, 742—Bulbo-capnine injected intravenously into dementia precox and phrenetic patients produced always an initial somnolence phase with a tendency for postural reflexes to become attenuated or to disappear. In the 2nd phase which is of longer duration and more characteristic and which might be called "the cataleptic phase," the postural reflexes are more or less exaggerated. PETER MASUCCI

The hypoglycemic action of "bean tea." G. SORGE *Boll soc ital biol ser* 5, 967-9(1930)—An aq. ext. of the pericarp of beans was made and administered to normal and diabetic subjects. 150 cc of an infusion from 20 g. of powd. pericarp lowered the blood sugar of fasting normals 15.5-38.6% in 1 hr. and 8.7-26% after 2 hrs. The same dose of infusion given to diabetics under analogous conditions caused a const. lowering of blood sugar, 12-29% after 1 hr. and 20.4-37% after 2 hrs. The same dose administered to normal subjects along with 50 g. glucose *per os* or $\frac{1}{4}$ hr. before the injection of 1 mg. adrenaline produced a const. and marked diminution of sugar and prevented the appearance of hyperglucemia. The expts. may be added to small doses of pancreatic hormone in the treatment of diabetics (synergic action). P. M.

The cure of intestinal lamblasis with naphthalene. G. SORGE *Boll soc ital biol ser* 5, 969-71(1930)—S. found naphthalene a satisfactory therapeutic agent in the treatment of *Lambia intestinalis*. The administration of 50 cg. per day in 3 capsules, one after each meal for 6-7 days, is recommended. PETER MASUCCI

The influence of acetylcholine on insulinemia. FILIPPO CONDORELLI *Boll soc ital biol ser* 5, 977-9(1930)—C. concludes that insulinemia diminishes after the injection of medium doses of acetylcholine. PETER MASUCCI

The inability of the bird organism to condense urea with glyceric acid. G. RUSSO-BONNI *Boll soc ital biol ser* 5, 980-1(1930)—Neither glyceric acid nor glycerol administered *per os* to chickens or geese caused the slightest increase in the amt. of uric acid excreted, even if simultaneous to the administration of glyceric acid, urea (1-1.5 g.) was injected into the pectoral muscles. The urea injected was found mostly in the excrements, while a small quantity was transformed into NH₃. Therefore, the bird organism does not have the capacity to transform urea into uric acid. P. M.

Relations between glucemia and phosphorus of the blood under the action of uranyl salts and extracts of thyme. MARIANO MESSINI AND ARISTIDE POLI *Boll soc ital biol ser* 5, 982-4(1930)—The injection intravenously of uranyl acetate (0.00012-0.00024 g. equiv. per kg. of rabbit) causes a diminution of the blood phosphates and simultaneously hyperglucemia without glucosuria. This is followed by a second hyperglucemic phase with glucosuria and anuria. The injection of thyme expts. is also followed by a diminution of blood phosphates and hyperglucemia. A lowering of the blood phosphates is apparently followed by an increase in the glucemic titer. P. M.

The behavior of the phosphates and glucose of the blood in normal dogs and in dogs deficient in adrenaline-producing tissue by the action of cold baths. IVAN CIACCIO *Boll soc ital biol ser* 5, 997-1001(1930)—The detns. were made on normal dogs and on dogs from which the suprarenal medulla had been removed. The animals were subjected to a cold bath, 5-7°, for 10-15 min. The results showed that hyperglucemia from cold bears no relation to the hyperfunction of the medulla or to the hypersecretion of adrenaline. There was not always a quant. agreement between glucemia and blood phosphates but in general hyperglucemia coincided with a marked lowering of inorg. P. Instead these showed a behavior analogous to that observed after the administration of glucose. PETER MASUCCI

The behavior of phosphageo and phosphates in muscle after cold baths. GRUSERRA SOLARINO *Boll soc ital biol sper* 5, 1001-3(1930) — Cold baths caused no variations in the phosphagen or phosphoric acid content of the striated muscles P. M.

The action of various chemical substances applied directly to the sensorimotor cortical centers of the dog. L. MORACCI *Boll soc ital biol sper* 5, 1003-6(1930) — The purpose was to see if any relation exists between the action of certain salts on the sensorimotor cortical centers and the position of the relative ions in the Hofmeister series. The salts studied were $(\text{NH}_4)_2\text{SO}_4$, NH_4Cl , NH_4NO_3 , NH_4CNS , K_2SO_4 , KCl , KNO_3 , KCNS , Na_2SO_4 , NaCl , NaNO_3 , NaCNS , Li_2SO_4 , LiCl and LiNO_3 . In another series LiClO_4 , LiI , Li_2CO_3 and Li lactate were studied. The results indicate that the action of the salts on the cortical centers is not a function of the pptg. power of the salts on colloids. With the exception of Li_2CO_3 , the Li salts exerted an exciting action; the exciting action of creatine was confirmed. There was an evident analogy between the motor symptoms provoked by Li salts and those induced by creatine. P. M.

The action of certain chemical substances on the motor activity of the intestine. R. DE MARCO *Boll soc ital biol sper* 5, 1006-6(1930) — The motor activity was tested by observing the length of time a small glass marble took to cross an intestinal loop under normal conditions and under the influence of adrenaline, atropine, pilocarpine and nicotine. Adrenaline and atropine prolonged the time, pilocarpine caused no changes, while nicotine shortened the time with respect to the normal conditions.

PETER MASUCCI

Adrenaline-like substances: their products of oxidation. PIETRO SACCARDI *Boll soc ital biol sper* 5, 1010-16(1930) — The substances studied were $(\text{HO})_2\text{C}_6\text{H}_3\text{COCH}_2\text{NH}_2$ (I) and $\text{HOCH}_2\text{CH}_2\text{CH}_2\text{NMe}_2$ (II). I was oxidized with Cl_2 water, dried on the water bath, extd. with alc., and the ext. evapd. to dryness (A), dissolved in H_2O (10% soln.) and injected subcutaneously into a guinea pig. The urine from the animal was dark, gave all the oxidation reactions analogous to the adrenaline blacks and gave a oeg. Thormählen reaction for melanuria, but the Ehrlich diazo reaction was pos. The urine was therefore alcaptonic and not melanotic. The alc.-insol. fraction (B) if treated with NaOH dissolved, and when injected produced the same phenomena of alcaptonuria. On analysis hydroxyhydroquinone was detected. II when oxidized and treated similar to I did not give rise to indole or pyrrole derivs. It was, therefore, not injected. The conclusion is drawn that oxidizability of these adrenaline-like substances is directly proportional to the no. of OH groups in the nucleus, and therefore, the more the nucleus is oxidized, the more it is oxidizable. The hypothesis is advanced that tyrosinase acts on adrenaline-like substances giving red and then black derivs. only if at least one H of the amino group is free.

PETER MASUCCI

Vagus function and adrenal insufficiency. GASTANO VIALE *Arch fisiol* 28, 9-24(1930) — In adrenal operated dogs, the vagus function is unchanged, the hemodynamic action of adrenaline is increased and that of choline is decreased. Vagotomy lowers the pressure and the rhythm of the heart. Preventive injections of ergotamine or ergotoxine inhibit the tachycardia caused by vagotomy in normal dogs. Ergotamine, given after vagotomy, causes a decrease of the heart rhythm. Ligation of the adrenals lowers the pressure and retards the rhythm of the heart after vagotomy. Tachycardia caused by vagotomy is due to adrenalinsmia. The sympathetic action on the heart is independent from adrenaline. Adrenaline probably exerts a central action on the vagus.

A. E. MEYER

The action of heavy metals on the spinal cord. MICHELE MITOLO *Arch fisiol* 28, 89-113(1930) — Cu , Ic , Au and Sn have a paralyzing effect on the coordinate sensory centers of the posterior cornu of the spinal cord, they are inactive on the motor centers. Mg , Hg and Pb paralyze the motor centers only. Zn produces depression in both centers. Ag , as AgNO_3 , has an exciting effect on the coordinate sensory centers but is inactive on the motor centers.

A. E. MEYER

Gomether as a general anesthetic. JOSÉ R. RÍOS. *Semana méd* (Buenos Aires) 1931, I, 396-7 — A mixt. of 80 cc. Et_2O and 10 cc. gometherol is recommended as an anesthetic in pulmonary lesions.

A. E. MEYER

Insulin in diabetes of children. PEDRO B. LANDABURU AND FÉLIX PUCULU. *Semana méd* (Buenos Aires) 1931, I, 446-50 — The daily doses must be divided into 2, 3 or even 4 injections in suitable relation to the meals.

A. E. MEYER

The local application of saturated solutions of magnesium sulfate. S. JOUAN. *Semana méd* (Buenos Aires) 1931, I, 459, cf. C. A. 25, 1009 — The effect on tumefactions, purulent inflammations, stings of insects etc., is due to osmotic changes. A summary of the beneficent actions is given.

A. E. MEYER

Cessation of attacks of auricular paroxysmal tachycardia by the use of calcium.

JOSEPH D. WOFFER AND HAMUR, DREIER. *Ann. Intern. Med.* 4, 705 (1911).

JOHN T. MYERS.

Tribromoethanol anesthesia. RALPH W. WATERS AND C. W. MURPHY. *Arch. Surgery* 21, 887 (1911). When kept at room temp., 1% solns of tribromoethanol are stable toward hydrolysis for 24-48 hrs. It has practically the same margin of safety as other commonly used anesthetics, i. e., $\frac{1}{2}$ of the fatal dose. With proper precautions, 0.1 g. per kg. body wt. can be safely given. It is useful, given intravenously, in very short operations, as it gives quick pleasant induction and recovery without gastrointestinal disturbance.

JOHN T. MYERS.

The toxicity of atmospheres containing hydrocyanic acid gas. JOSEPH HANCIK. *J. Hyg.* 31, 1-31 (1911). The sp. power of HCN gas to increase the total ventilation is an important factor in determining the toxicity of the gas. Man is relatively insensitive to the inhalation of HCN. The time of exposure sufficient for collapse is about one third of the fatal period of exposure. The extreme susceptibility of the canary makes it very valuable as an indicator of lethal concentration where a chem. indicator is not available. Pigeons are as good since they succumb at the lethal concentration for the canary, although their lethal concentration is about twice as great. Treatment consists of artificial respiration, and the administration of glucose and nitrates.

JOHN T. MYERS.

Experimental dust inhalation in guinea pigs. P. HAYNES. *J. Hyg.* 31, 95-121 (1911). Most inhaled particles remain on matter at least to a small extent. The solute may be harmlessly active or toxic. If harmless the cell is stimulated to detach itself from the alveolar wall and to remove the dust. If toxic the solute affects the viability of the phagocyte which becomes less able to detach itself. At the same time the solute diffuses into the neighboring tissues with irritation to them and subsequent ulcers. The more and longer a substance causes the greater damage, hence the solute has much to do with extent of damage. $\text{Al}(\text{OH})_3$ is the dust par excellence in predisposing to tuberculosis. The harmful effects of $\text{Al}(\text{OH})_3$ may be neutralized by simultaneous administration of basic dusts as $\text{Al}(\text{OH})_3$ or MgCO_3 although the latter are harmful when inhaled alone. It is suggested that the respective solutes combine to form non-toxicates which are harmless.

JOHN T. MYERS.

Asbestosoid bodies in the sputum: a study of specimens from fifty workers in an asbestos mill. P. W. SIMON AND A. SITTIKIAN AND SIRACHIAN. *J. Path. Bact.* 34, 1-4 (1911). The inhalation of asbestos dust in high concentration leads to the appearance of asbestosoid bodies in 48 out of 50 workers examined. The sputum is always mouldy in character. It may resemble egg albumin. The bodies seem to be the result of degeneration of the asbestos fibers or may be some substance eliminated by the cell. Nontubercle bacilli were found in any of these cases.

JOHN T. MYERS.

Experimental liver necrosis from shale oil. CYRIL J. PEARSON. *J. Path. Bact.* 34, 5-12 (1911). Liver necrosis was produced by tablets and also by the intraperitoneal injection of shale oil. By graduating the doses it is possible to produce all stages from acute necrosis to cirrhosis. The poisonous action of the non-exchangeable residue of shale oil is very slight. The liver was not thus affected in other animals. J. T. M.

The effect of oils on gastric secretion and motility. W. MINOR, R. RUMMEN. *Quart. J. Med.* 24, 131-52 (1911). The efficiency of oils in inhibiting gastric secretion appears to bear some relationship to the degree of saponification of the constituent fatty acids, less saponifiable being more efficient. The relationship between efficiency and sp. gr. and viscosity is less significant. Of related substances, only free fatty acids show inhibitory properties. Oil diminishes the response of the stomach to food or other substances but not to the effect of histamine. Oil inhibits both physical and chemical stimulation of secretions. Its effect ceases almost entirely after it leaves the stomach. It diminishes the motility of achlorhydria. The effect on gastric motility is not secondary to its effect on secretion. It seems necessary to assume that oil after absorption exerts some restraining influence on the gastric cells and the motility, or that it leads to the formation of the liberation of some substance which acts in a similar manner.

JOHN T. MYERS.

Cholagog action of a few derivatives of the aliphatic series of low molecular weight. HENRIK CHABROU, R. CHABRONAT, M. MAXIMIN AND R. WATZ. *Compt. rend. Soc. Biol.* 106, 15-6 (1911). A study was made of the effects on the cholagog action caused by substituting various groups and atoms for the Cl and the CO_2H group in $\text{CH}_3\text{CO}_2\text{H}$. Conclusions: (1) The cholagog action disappears when the CO_2H group is replaced by an amide, etc. or Me group. (2) $\text{ArCH}_2\text{CO}_2\text{H}$ causes an immediate and powerful cholagog action but its toxicity is too great for a prolonged study; glycine in a 2 gr. dose doubled the biliary flow during 1 hr. with a 17.5 kg. dog but was completely inactive in doses of 1 gr. with other dogs; glycolic, cyanacetic, glyoxylic and malonic acids had no

effect in 1 gr doses with dogs of about 20 kg; hippuric acid had a very distinct action on 2 animals. (3) Of the compds in which both groups are replaced at the same time, EtCl is completely inactive but FIBr or EtI increases the chologag action when large doses are used. CHCl_3 and CCl_4 are completely inactive. ALBERT L. RAWLINS

The chologag action of chlorine derivatives of acetic acid. ESTEVE CHABROL, R. CHASOVNAT, M. MAXIMIN AND R. WATTE. *Compt rend soc. biol* 106, 17-8(1931). — There is no close relationship between the chologag action and the no. of Cl atoms in the Cl derivs of AcOH . AcOH , $\text{ClCH}_2\text{CO}_2\text{H}$, $\text{Cl}_2\text{CHCO}_2\text{H}$ and $\text{Cl}_3\text{CCO}_2\text{H}$ were tested in molecularly equiv. doses of their Na salts, the acids represented thus amounted to 0.40, 0.60, 0.80 and 1.0 gr resp. AcOH had no chologag action, $\text{ClCH}_2\text{CO}_2\text{H}$ gave const. results, at least doubling and sometimes quadrupling the flow of bile and at the same time increasing the concn. $\text{Cl}_2\text{CHCO}_2\text{H}$ had no action; $\text{Cl}_3\text{CCO}_2\text{H}$ doubled the biliary flow and increased the concn. of bile from 2.70 to 3.15. ALBERT L. RAWLINS

Respiratory stimulation by sodium sulfide. C. HEYMANS, JEAN J. BOUCKAERT AND L. DAUTERBANDE. *Compt rend soc. biol* 106, 32-4(1931). — Two series of expts. with dogs show that the stimulating action of Na_2S on respiration is largely reflex from the carotid sinus and is only very slightly due to a direct stimulating effect on the respiratory center. ALBERT L. RAWLINS

Carotid sinus and respiratory reactions by cyanide. C. HEYMANS, JEAN J. BOUCKAERT AND L. DAUTERBANDE. *Compt rend soc. biol* 106, 54-5(1931). — A series of expts. was performed with dogs in which the cyanide was caused to act solely on the carotid sinus and then solely on the centers. The results show that the inhibitory action of cyanide on respiration is due to the sensitiveness of the carotid sinus to cyanide, and not due to the action of the cyanide on the centers. ALBERT L. RAWLINS

The hypoglycemic action of sulfur. Sulfur mineral waters. I. DANIEL AND M. POPESCU BUREU. *Compt rend soc. biol* 106, 109-7(1931). — The action of sulfur water on dogs and normal and diabetic patients is contrasted with the action of distd. water. In all cases where the sulfur mineral water was used there was a pronounced diminution of blood sugar, while distd. water had no effect. The authors believe that the hypoglycemic action should be attributed to the colloidal S and H_2S in the water. ALBERT L. RAWLINS

The action of commercial glucose, comparatively administered by the digestive and intravenous methods, on the glycogen content of the liver, heart and muscles. D. SPICER, M. POPESCU AND A. CRAIPALEANU. *Compt rend soc. biol* 106, 112-5(1931). — The glycogen content of the liver, heart and muscles was detd. with 12 rabbits administered com. glucose by digestive and intravenous methods following a fast. To increase the liver glycogen the digestive method is more effective than the intravenous while to increase the glycogen content of the muscles, and particularly of the heart, the intravenous method is incomparably superior. ALBERT L. RAWLINS

Pharmacodynamic action of hydrogen sulfide introduced into the digestive tract. F. KMIETOWICZ. *Compt rend soc. biol* 106, 126-8(1931). — In a study with dogs the following behavior of H_2S was demonstrated: large doses introduced into the intestine first cause an acceleration of respiration and then cessation with death resulting; the blood pressure is raised, the heart rate is lowered, with an increase of amplitude; artificial alkalosis, acidosis and the injection of morphine increase the toxicity of H_2S , alkalosis having the greatest effect. The mechanism of the above actions is discussed. The absorption of H_2S is greater in the small intestine than in the large intestine. H_2S shows analogous but less distinct behavior when introduced by mineral water. ALBERT L. RAWLINS

Action of iodine, calcium and magnesium ions on the oscillographic index and arterial pressure in transcranial di-electrolysis. GEORGES BOURGUIGNON AND SOCRATE ELIOPOLLOS. *Compt rend.* 192, 379-82(1931). *Compt rend soc. biol* 106, 454-8(1931), cf G. Bourguignon, *Rev. Actinologie* 3, 11, 180, March-April 1929. ALBERT L. RAWLINS

"Dial" as a surgical anesthetic for neurological operations with observations on the nature of its action. J. F. FULTON, E. G. T. LIDDELL AND D. McK. RIOCH. *J. Pharmacol.* 40, 423-32(1930). — For cats and dogs 0.5 cc. per kg. administered intraperitoneally was sufficient for surgical anesthesia. The motor cortex still reacted to elec. stimulation. From this fact, and from other observations (of chop licking, salivation, vomiting, indicating an effect on the medulla, and of absence of excitement and a resemblance to natural sleep) it is suggested the site of action of Dial is in the hypothalamus and medulla. C. RIEGEL

The calorogenic action of morphine as revealed by addiction studies. H. G. BARBOUR, D. E. GRECO AND L. G. HUYER. *J. Pharmacol.* 40, 433-50(1930). — Three

of 4 dogs after 7-10 weeks on morphine showed a basal metabolism higher than normal, and this high value dropped on withdrawal of the drug. Basal N excretion was decreased, although total N output increased. Carbohydrate combustion increased, fat combustion decreased. It is suggested that narcotic action may be due to inhibitory impulses arising peripherally from the calorigenic stimulation. C. RIEGEL.

The reaction of iodates in vivo. L. C. MAXWELL. *J. Pharmacol.* 40, 451-5 (1930).—Lethal intravenous dose of NaIO_3 was 100 mg per kg for rabbits. A dose of 75 mg caused appearance of toxic symptoms (convulsions, spastic paralysis, irritation of mucus membranes, albuminuria). The toxicity of the iodate increased in the presence of NaI . The toxicity apparently is not due to liberation of I in the tissues. C. R.

Urinary sulfur and thiocyanate excretion in cyanide poisoning. RALPH G. SMITH AND RUSSELL L. MALCOLM. *J. Pharmacol.* 40, 457-71 (1930).—Rabbits on a const. diet were allowed to breathe HCN fumes for certain periods for a number of days. Inorg. and conjugated sulfates, neutral S and thiocyanate excretion in urine were detd. Neutral S increased during cyanide administration, the increase being accounted for by excretion of thiocyanate. Inorg. sulfate decreased in most cases, so that total S was the same, or less. An increase in total N caused an increase in N:S ratio. After intravenous injection of NaCN 61-100% of the cyanide given was recovered in the urine during the 1st three days as thiocyanate. Conclusions: Increased urinary neutral S is due to detoxication of cyanide rather than depressed oxidation processes. Thiocyanate formation is the principal method of cyanide detoxication. C. R.

Hemoglobin production. II. The relief of anemia, due to milk diet, by feeding amino acids. DAVID L. DRABKIN AND H. K. MILLER. *J. Biol. Chem.* 90, 531-43 (1931).—The effect of the addn. of amino acids relatively free from Fe to anemic rats on a whole-milk diet contg. 0.3 mg Fe per day was studied. All the preps were free from Cu and from other heavy metals. Alanine, alanine-HCl, histidine-2HCl and HCl alone had no effect on hemoglobin production, indicating that in this type of anemia the relief of achlorhydria is not essential (cf. McLester, *J. Am. Med. Assoc.* 95, 719 (1930)). Arginine and Na glutamate rapidly produce continuous hemoglobin regeneration. Tryptophan, Na aspartate, pyrrolidonecarboxylic acid and a proline mixt. from hydrolyzed gelatin produce a temporary increase in hemoglobin production. It is suggested that the effective substances act as a source for the synthesis of pyrrole rings for the heme mol.; the absence of either Fe or of these substances produces anemia. K. V. TIMMANN.

A study of electrolyte equilibrium in the blood in experimental acidosis. HENRY N. HARKINS AND A. BAIRD HASTINGS. *J. Biol. Chem.* 90, 565-95 (1931).—The effect of injecting 2-3 cc N HCl per kg body wt. into dogs was followed by CO_2 , p_{H} , ion distribution and refractive index measurements. The acid caused an anuria so that very little of the excess Cl was excreted. The p_{H} fell to 6.0 and in 2 cases to 6.1. Hyperventilation began 20 sec. after the start of the injection, but respiration became normal again in 1 min., while the p_{H} had not returned to normal in 3 hrs. or more. The respiratory center was not inactive, as shown by response to maximal elec. stimulus, and hence the p_{H} is not a respiratory hormone. The distribution of HCO_3^- and Cl ions between cells and plasma was according to the Donnan relation throughout (Van Slyke, Wu and McLean, *C. A.* 17, 3046), even on the acid side of the isoelectric point of hemoglobin. The refractive index of the serum was increased, and the water content decreased. The water content of the cells, however, and the Cl contents of both cells and serum were increased. Two expts. on blood *in vitro* gave essentially similar results. K. V. T.

The comparative effect of strychnine in the normal and dehepatized frog. J. T. PRIESTLEY. *Proc. Staff Meetings Mayo Clin.* 6, 71 (1931).—Liverless frogs showed a reaction to a given dose of strychnine that was more than twice as severe as that given by normal or splenectomized frogs. Partially dehepatized frogs were distinctly more susceptible to strychnine than normal or splenectomized frogs. R. C. WILSON.

β -Phenylethylamines. I. Mescaline and mescaline-like substances (SLOTTA, HELLER) 10. Some synthetic bases similar to ephedrine (CHERBULIEZ, et al.) 10.

JOHANNESSEN, FRITZ. Kinine in de geneeskundige Prachtijk aan de Hand van de Gegevens der Pharmacologie. Amsterdam: Bureau tot Bevordering van het Kinine-Gebruik. 210 pp.

12—FOODS

F. C. BLANCE AND H. A. LUPPER

An outbreak of food poisoning in Staffordshire. I. The clinical manifestations. E. S. CLAYTON AND V. E. MILNE *Brit Med J* 1930, 11, 684-5. II. Bacteriological investigation. JOHN MENTON *Ibid* 685-6.—A report is given of a limited outbreak of food poisoning due to eating pie made from pork infected with hog cholera bacillus.

J. B. BROWN

Use of buffers in determination of color by means of titanium trichloride. II. O. L. EVENSON AND R. H. NAGEL *Ind Eng Chem, Anal Ed* 3, 167-9 (1931), cf. C. A. 22, 3780.—Analysis of food colors by titration with $TiCl_3$ is made fairly accurate by the addn. of properly chosen buffer catalysts such as Na citrate or bitartrate.

H. F. LAUTOLD

The viscosities of alkali solutions of pure wheat flours and their relation to practical baking research. THEODOR BIECHY WITH ROBERT MÜLLER, GRAMS AND NÜCHTER *Kolloid Z.* 54, 194-210 (1931).—The viscosities of 0.05 N NaOH solns. of wheat flours are const. after 2 hrs. and were detd. at 20° with a Lawaczeck falling body viscometer. A standard comparison may be made by comparing the viscosity of a soln. in 0.05 N NaOH contg. enough flour to contain 1 g. of gluten per 100 cc. The gas-holding capacity of dough was detd. at 32° with the Tullmann app. The following relation was found to hold for the wheat flours tested [$Teiggarkraft \times \text{standard viscosity} / \text{Bread vol. (100 g. flour)} = \text{const.}$, where $Teiggarkraft$ is the difference between original and final dough vol.].

ARTHUR FLEISCHER

Bakers' yeast. I. Influence of age upon the fermentation time of this yeast. E. ROSENBAUM *Z. Untersuch. Lebensmittel* 59, 607-12 (1930).—The manuf. of yeast cake from beet-sugar molasses is described. Analytical data on 20 samples of yeast are given. The moisture contents as detd. by heating 4 hrs. at 105-8° were 72-3%. High moisture content is associated with poor keeping quality. Krepung quality was detd. by the time required for the cake to soften when heated at 35°, the times varied from 55 to 105 hrs. The N contents varied from 1.81 to 2.52% of the dry substance, corresponding to protein contents of 41.38-57.25%. The fermentation time is the time taken by a dough of standard compn. pressed into a mold of fixed dimensions to rise 5 times in succession to a height of 7 cm. above the mold. The sum of the 5 periods, or total fermentation time, showed great variation, 176-271 min. The fermentative power of the yeast decreases rapidly on storage, but deterioration is reduced by storage at low temps.

C. R. FELLERS

Development and heat economy of the gas bakery oven. W. RAISS *Arch. Wärmerw.* 11, 381-6 (1930).—The development of the various German types of gas ovens for baking bread is discussed. A successful gas oven will not imitate the models for solid fuels, but must have a min. of wt. and good insulation.

E. W. T.

Milk considered as a complete and biologically balanced food. RAOUX LECOQ *Bull. soc. hyg. aliment.* 18, 481-509 (1930).—A review.

A. PAPINBAU-COUTURA

The pigments of milk. MARTIN LUNNBOG *Biochem. Z.* 231, 274-89 (1931).—The natural yellow color of butter fat is due to carotene, as can be ascertained by the carotene spectrum of the unchanged fat. The tintometric detn. of the carotene according to Lovibond is unreliable because of the very min. quantities present. More reliable results are obtained from studies on the sapon residue. Five cc. of a freshly prepd. 20% alc. KOH is used for the sapon, per g. fat, and the sapon is carried out for 45 min. over a boiling water bath in an Erlenmeyer flask provided with a long reflux tube. The reaction mixt. is dil'd with 10 vols. of H_2O , and this is then extd. 4 times with $1/4$ its vol. of ether. The combined ether exts. are dried 2 hrs. with anhyd. Na_2SO_4 and evapd. at 50-55°, the last stages *in vacuo*. The residue is transferred quantitatively to small measuring cylinders with small quantities of anhyd. chloroform, and the tintometric detns. are made on these solns. For this purpose the chloroform soln. is layered over $SbCl_3$ soln., and the contents are carefully mixed and exam'd in the tintometer. Direct extn. of the milk fat is impossible, the extn. of the fat with petr. ether from milk acidified with lactic or acetic acid is also full of difficulties, since the acid destroys the carotene when the petr. ether is evapd.

S. MORGILLIS

Determination of milk fat. A. C. ANDERSON *Z. Untersuch. Lebensmittel* 59, 600-2 (1930).—In the detn. of fat in milk which is in an incipient state of decomp., sepn. of the fat as in the Rose-Gottlieb method may lead to low results because of sapon., even if $K_2Cr_2O_7$ has been added as a preservative. It is preferable to use the Schmid-Bondzynski-Ratzlaff method as follows. Ten g. of milk is mixed with 10 cc. of HCl .

d. 1.10, and boiled for 4-5 min., heat is applied as gently as possible to avoid discoloration. After cooling, the mixt is extd with 10 cc of alc, 25 cc of Et_2O and 25 cc of light petroleum as in the Rose Gottlieb method. C. R. FELLERS

Determination of the milk proteins. III. Proposed modified method for casein. GEORGE M MOIR *Analyst* 56, 147-9(1931), cf. C. A. 25, 1597.—The following modified method is proposed for the detn of casein. Dil 10 ml of milk with 50 ml of water which has been heated to 40-42°. Add at once 15 ml of 1.67 N AcOH and stir by rotating the contents of the beaker 4 times against a stirring rod. After 20 min add 4.5 ml of 0.25 N NaOAc soln and filter after it has stood an hr or more. Wash promptly and det the N content of the casein by the Kjeldahl method with Na alizarine sulfonate as indicator in the final titration. The directions are given in detail together with several important explanatory notes. W. T. H.

The determination of the calcium, magnesium and acid-soluble phosphorus of milk by means of trichloroacetic acid filtrates. GEORGE P. SANDERS *J Biol Chem* 90, 747-50(1931).—The proteins of milk can be completely pptd by mixing 4 vols of 10% $\text{CCl}_3\text{CO}_2\text{H}$ with 1 vol of milk. The filtrate can be used for the detn of Ca and Mg with a high degree of accuracy, and 1 working day is saved in the prepn. of samples. The procedure should also be of value in studying P combinations in milk with particular reference to casein and lipid P. In several milk samples 68.4 to 81.9% of the P was found to be acid sol. There was no definite correlation between the amt of acid insol. P and the casein content, but in general milk high in casein was high in acid-insol P. A. P. LOTHROP

Determination of the volatile fatty acids in the milks of the cow, goat and sheep. VITTORIO BORRA *Pediatrics rivista* 39, 102-9(1931).—In the course of his investigations on the origin of the anemia resulting from the use of goat milk, B made comparative analyses on the quantity of volatile fatty acids present in the milks of cows, sheep and goats. The Reichert-Meisli no. for the milks of the goat, cow and sheep was 7.72, 3.68 and 2.53, resp., while the Polenske no. was 3.76, 1.49 and 0.87, resp. G. S.

Simple calculation of the limit of value of the microscopical examination of milk for tubercle bacilli. D. R. WOOD *Analyst* 56, 179-80(1931); cf. C. A. 24, 5337.—Without concn. (and the concn. methods are not really successful) from 10 to 500 bacilli per cc. can be detected by microscopical examn. If cases are excluded in which there is a bulky deposit, characteristic of tubercular mastitis complicated by streptococcal mastitis, and if the attention is confined to milk giving a normal deposit of not more than 2.5 cc. in a vol of 100 cc., the microscopic detection is limited to the detection of not less than 10-50 bacilli per cc. and is, therefore, less than 1/500 as sensitive as the biological method. W. T. H.

Crystallization of lactose in sweetened condensed milk. B. SROAL. *J. S African Chem. Inst* 13, 33-40(1930).—The texture of sweetened condensed milk is a function of the size of the lactose crystals. In the cooling of condensed milk during manuf. the immediate production of a large no. of crystal nuclei, rather than the growth of crystals around a small no. of centers, should be the aim. B. C. A.

Studies on the use of dry skim milk in manufacturing ice cream. C. D. DAILEY AND J. I. KEITH *Ice Cream Trade J.* 27, No. 1, 33-5(1931).—When dry skim milk of good quality was used in place of condensed skim milk in ice cream mixes contg. cream and milk, the resulting ice cream scored as high as the condensed skim milk controls. When dry skim milk was used in conjunction with butter, the score of the ice cream was lower than when cream was used. Mixes contg. butter and skim milk powder yielded the desired overrun more slowly than mixes contg. cream and powder. From the standpoint of flavor, texture and body dry skim milk of good quality can be successfully used as a substitute for condensed skim milk in mixes contg. cream and fluid milk as a source of fat. One half lb. of dry skim milk can be successfully used in one gallon ice cream mix under the conditions mentioned. A. H. JOINSON

The causes and prevention of cream feathering. F. J. DOAN. *Creamery and Milk Plant Monthly* 20, No. 1, 33-8(1931).—Fresh unhomogenized table cream rarely exhibits feathering in coffee. Creams which have become acid on account of fermentation and homogenized creams are susceptible to feathering. The feathering of homogenized creams was shown to be due to fat clumping. Homogenization of cream in such a way as to prevent fat clumping reduced the tendency to feather. Cream for prepg. homogenized table cream should be fresh, of low acidity, and should be pasteurized at 160-62°F. for 30 min. or, if a flash method is used, at a temp. of 175°F. The cream should then be homogenized at the same temp. If single-stage homogenization is used, the cream should be rerun at once while hot at about 1000 lb. pressure. When a two-stage machine is used, the pressure on the second valve should not be in

excess of the pressure on the first. If leathering cannot be prevented by these procedures certain salts such as Na citrate, NaHCO_3 , and Na_2HPO_4 , may be added. The amt. of these salts to be used varies but 0.025 to 0.10% will usually be sufficient. The salt should be added prior to homogenization. A. H. JOHNSON

The hygienic supervision of butter. A. D'AMMASSIO *Ind. olii minerali e grass* 10, 169-71 (1930) of *C. A. 24*, 5888.—On 114 samples of butter taken between December and April when the cows were fed with dry hay, only 8 had the value of the sol. volatile acids under 26. Feed, such as grass, richer in volatile acids, produced butter richer in these. The stage of lactation has an effect on the production of rich milks. After the period of the colostrum, the quantity of the volatile fatty acids decreases from a very high figure to normal. Other conditions, such as estrus, weaning of the calf, etc., affect the compn. of the milk. Many controlled samples of butter gave the const. limits. Zeiss refractometer reading at 35° 43.2-49.5, sol. volatile acids no. 20.8-35.15, insol. volatile acids no. 11-30. R. SANSONE

New methods of analysis of oils and fats, with particular reference to butter analysis. S. FACHINI AND G. DORTA *Atti III congresso naz. chim. pura applicata* 1930, 683-7.—Three and a half g. of butter is melted and filtered, then the sapon. no., sol. fatty acids, fixed fatty acids, I no., mol. wt. and n of fixed fatty acids are detd., and the fixed fatty acids are sep'd into a group contg. caprylic, capric and oleic (linoleic) acids not pptd. from soln. in acetone by KOH and a group contg. lauric, myristic, palmitic and stearic acids pptd. from soln. in acetone by KOH. The I no., mol. wt. and n of the first group are calcd. From the I no. the content of caprylic and capric acids and the av. mol. wt. of these 2 acids in the mixt. can be calcd. The n of the first group, for all grad's of butter tested, was 1.4510-1.4500 at 25° , palm oil gives 1.4391-1.4380. E. M. S.

Analysis and composition of vegetable parchment used for packing dairy products. PAUL ARUP *Analyst* 56, 149-61 (1931).—The following detns. were carried out on 41 com. samples of parchment paper, water, ash, water sol. ext., reducing material such as dextrose, bursting strength (wet and dry), wt. per ream and power to support mold growth. The methods used are described and the results tabulated. As a result of these detns. the conclusions drawn are: The moisture content should not exceed 10%. The bursting strength, as detd. by the Ashcroft tester, should be at least 25 lb. per sq. in. for parchments weighing 25 lb. per ream and 18 lb. per sq. in. for material weighing 18 lb. per ream. The strength immediately after being wetted should be at least 33% of the strength of the unwetted material. The ash should not exceed 0.45% and the water sol. ext. 1.30%. The percentages of reducers in the water sol. ext. were calcd. as a guide for the detection of added sugar. Certain standards adopted by Germany (8% of ext. and 3% of ash) are considered too high, being due probably to the addn. of sugar and salt. An unduly high content of ext. renders the paper particularly susceptible to attack by mold. The water sol. ext. from genuine parchment consists of the decompn. products of lignocellulose and can be characterized by methods of Cross and Bevan. W. T. H.

Fatty acids and component glycerides of Indian ghee. R. BHATTACHARYA AND T. P. HILDITCH *Analyst* 56, 161-70 (1931).—Ghee is clarified butter or cream used in India in place of butter which, because of climatic conditions, is not suited to Indian methods of cooking. To prep. ghee, the thick scum that forms on boiled milk when it is allowed to stand in a warm place is skimmed off, mashed and churned with added cold water toward the end. The resulting buttermilk is removed, and the residual fat is heated gently in an open pan, strained and stored. Two samples of cow ghee and 2 of buffalo ghee were analyzed, one sample from a stall fed animal and the other from a pastured one in each case. The butter fat content of the cow ghee was similar to that of ordinary butter. The Reichert Meissl values and the % of butyric acid were toward the lower values obtained by Hilditch and Sleightholme in butter from New Zealand cows, but the oleic acid content was normal. The stearic acid content is higher in the fat from the stall fed cow than from the pastured animal, and the palmitic acid value is normal for butter fat or beef tallow. The corresponding specimens of buffalo ghee were not very different, although there is distinctly more butyric acid which corresponds to somewhat higher Reichert Meissl and Kirschner values. The most apparent difference lies in the increased content of stearic acid and the presence of arachidic acid. Broadly speaking, it would appear that the main characteristics of the fatty acids in a fat are fairly well defined for a given species of animal, subject to minor variations due to varying climatic and feeding conditions. With respect to the glyceride structure of the ghee, it is of the mixed heterogeneous type. There is no evidence of segregation of any fatty acid into a simple triglyceride, and all the said fatty acids are found, in more or less the same ratio, in each part of the fat. It

is doubtful whether triolein is present in appreciable amt., and the tripalmitin or tristearin present must be small if any. The 4 fats studied were of first quality, comparable in flavor and odor to the best butter fats. W. T. H.

The botany and chemistry of the soy bean. W. RIEDE AND B. REWALD. *Landw. Vers. St.* 110, 291-304 (1930).—Seven samples of soy bean raised in Germany analyzed N 5.97, 6.62, protein 36.59-41.37, fat 11.67-17.39 and phosphatides 1.02-2.53%. These figures are reduced to a basis of 12% H₂O. JOHN R. HILL.

Determination of moisture in soy beans. L. JOLSON. *Mém. Univ. Est. Extrême-Orient* 13, No. 6, 1-23 (1929). *Chimie et Industrie* 25, 145 (1931).—From a study of drying at 100-5° and at 120-40°, of distn. in presence of xylene, and of the effects of the fineness of grinding, wt. of samples and temp. and time of drying J. concludes that all 3 methods give satisfactory results provided the conditions are properly selected and strictly adhered to. He recommends grinding to pass a 15-25-mesh sieve and drying 5 g. for 20 min. at 130°. Drying at 100-5° requires 5-8 hrs. and should not be continued to const. wt. A. PARINCAU COUTURE.

Mineral, nitrogen and fat content of some varieties of mature bean seed and of string beans. J. S. MELLARGUE AND W. R. ROY. *J. Nutrition* 3, 470-81 (1931).—Twelve samples were analyzed for protein, fat, etc., as well as for mineral content. Both string beans and ripe beans are very rich in mineral constituents and S. It is probable that the curative property of beans for beri beri is associated with their relatively high mineral content. C. R. FELLERS.

Carbohydrates of the lotus rhizome. TAIRO TAKAHASHI AND HATSUNAO YOKOYAMA. *J. Agr. Chem. Soc. Japan* 6, 870-2 (1930).—The lotus rhizome is a common food in Japan. Most of the sugar fraction was non-reducing sugar. From the hot ale ext. of the rhizome powder sucrose was isolated. Raffinose, galactose and galactane could not be detected. The polysaccharides were mostly starch and cellulose. Y. K.

Vitamin B content of raw and canned beef. YOSHIO UMETSU. *Imp. Zootech. Rpt. Sta. (Japan)*, *Bull.* 25, 12 pp. (1930).—1 or normal growth of rats, the min. daily requirement of vitamin B is supplied by an ale ext. of 20 g. raw or 40 g. canned beef. The vitamin B content of canned beef is not diminished during storage for 4-5 years. K. KITAHARA.

Detection of the onset of decomposition in meat by the content of ammonia in the form of salts. P. M. LITTELSCHNEIDER. *Z. Untersuch. Lebensmittel* 59, 599-600 (1930).—The effect of NH₃ gas from a leaky NH₃ system for cold storage on the appearance and quality of meat is described. NH₃ contamination may be detected by testing the rinsing water with Nessler's soln. In conducting the Glissmann-Rochwarger test (C. A. 24, 3083) for incipient putrefaction and for the detection of NH₃ in flesh, L. recommends that the outer and inner portions of the meat be examd. separately to det. whether the meat has been exposed to NH₃ gas or has been dressed with NH₃ salts as preservatives. C. R. FELLERS.

Citrus fruit production in the Lower Rio Grande Valley of Texas. HAMILTON P. TRAUB AND W. H. FAHND. *Texas Agr. Expt. Sta. Bull.* 419, 60 pp. (1930).—Very complete pomological, phys. and chem. data for grapefruit, oranges, lemons, kumquats, tangerines, etc., are given. For grapefruit, total sol. solids range from 8.3 to 10.5%, total sugars from 5.5 to 7.0, total acidity as citric from 0.96 to 1.4, fat from 3.1 to 3.3, protein from 0.42 to 0.53 and ash from 0.19 to 3.0. C. R. FELLERS.

Acid to sugar ratio in oranges. D. J. R. VAN WIJKE. *J. S. African Chem. Ind.* 13, 52-8 (1930).—Examination of 3 typical varieties of oranges shows that maturity as commonly judged from the ratio, acid/degrees Brix of the juice, is unsatisfactory. A better basis of comparison is the ratio acid/"caked sugar" (where caked sugar = degrees Brix - acid - 1.75). The value 1.75 represents the contribution of extraneous matter, other than acid and sugar, to the d of the juice and is fairly const. for all samples examd. B. C. A.

Banana—a challenge to chemical investigation. H. VON LORSECKER. *J. Chem. Educat.* 7, 1537-43 (1930).—Bananas soften on ripening because of the change of protopectin into pectin. When overripe the pectin has changed to acid. Ethylene has little effect on ripening bananas except on "dormant fruit." Banana vinegar compares favorably with cid. r. vinegar. Ripe banana flour has some therapeutic value. Bananas are rich in vitamin A and C and also contain B, G and E. ANN NICHOLSON HIRD.

Effect of solid carbon dioxide upon transportation diseases. CHARLES BROOKS. *Phytopathology* 21, 103 (1931).—It is possible by the use of solid CO₂ to secure a CO₂ content of the atm. of a car within 30-60 min. which will check rotting and softening of warm fruit as much as would a 30-40° drop in temp. It is essential that the gas escape within the next 18-21 hrs. or injury to flavor may result, especially with peaches.

red raspberries and strawberries. Dewberries, blackberries, cherries and plums are more resistant to injury. Grapes, sweet corn, peas and beans offer the greatest promise of beneficial effects without harm to the product. JOSEPH S. CALDWELL

Photoactivity of honey. J. SITZ. *Z. Untersuch. Lebensmittel.* 59, 606-7(1931); cf. C. A. 23, 2223.—The possibility of Ra in the ash of honey was investigated by placing the ash as well as natural honey in both glass and quartz containers and subjecting them for 30 days to contact with photographic plates. Other samples were irradiated with ultra violet light and Röntgen rays and similarly exposed. All results were negative. C. R. FELLERS

A rapid method for the determination of crude fiber. U. H. PURANEN AND E. S. TOMILA. *Acta Chem. Fennica* 3, 85-9(1930).—The Weende method consisted in treating 3 g. of the crude fiber with acid to remove fats, boiling for 1/2 hour with 200 cc. of 1.25% H_2SO_4 , filtering and washing until wash water was neutral. The residue was boiled for 1/2 hour with 1.25% KOH, washed, dried and ignited. The process took several days due to the colloidal nature of the soln. obtained and the difficulty in filtration. If the vol. fell below 172-181 cc. on boiling, the analysis showed 23.64% crude fiber (theoretical 26.08%), while if the vol. was 223-227 cc., the result was 26.53%. A later method involved the direct neutralization of the acid with KOH so that a large excess of KOH was present during the second boiling. The results were always high. The proposed method was to reflux 3 g. of finely ground fiber with 50 cc. of 5% H_2SO_4 and 150 cc. of H_2O for 1/2 hr. The flask was allowed to cool somewhat and 20 cc. of 28% KOH was added with a pipet. After refluxing 1/2 hr., cooling, filtering on asbestos, washing with hot H_2O , with 1.25% acid, again with hot H_2O , and finally with alcohol and ether or acetone, the residue was dried at 105° and ignited. S. A. KASJALA

New procedure in the chemical investigation of fodder and foodstuffs. J. KÖNIG, F. BARTSCHAT AND B. STEMPER. *Z. Untersuch. Lebensmittel.* 59, 164-72(1930).—The deficiencies in the normal methods of recording the constituents of foodstuffs are enumerated, and a systematic scheme is presented for analysis and separation into the following groups: water content, fat (Et_2O ext.), crude protein (amide + pure protein, both H_2O -sol., and sol. or insol. in 2% HCl), water sol. carbohydrates (acid as lactic acid + sugar reducing powers before and after inversion + dextrin + residue), starch, proto-cellulose and hemicellulose (including hemicelluloses); α -cellulose and β -lignin including α -pentosans, insol. lignin, cutin and suberin, mineral matter (H_2O -sol., and sol. or insol. in 2% HCl). The results of analyses of wheat, flour and bran are tabulated as examples. C. R. FELLERS

A new method for the "sand determination" in the presence of silicic acid in feed stuffs. W. LEPPER. *Landw. Vers. Sta.* 110, 305-11(1930).—Five g. of the substance is ashed over a small flame. Twenty cc. NaOH(15%) is added to the ash in a Pt or porcelain dish which is covered with a watch glass and heated on a water bath for 1/2 hour. It is then washed into a beaker, diluted to 400 cc., and 100 cc. HCl (1:1) is added to it. The mixt. is stirred and filtered and the ashed residue is the "sand". JOHN R. HILL

Report on the work of the laboratories of the Central Experiment Station on feeding. I. Bioenergetics laboratory. J. LEFÈVRE AND A. AUGUET. *Bull. soc. Hyg. Aliment.* 18, 328-39(1930). II. Biological chemistry laboratory. M. JAVILLIER. *Ibid.* 339-43. III. Physiological laboratory. L. RANDON. *Ibid.* 369-89. IV. Laboratory and experiment station for investigations on livestock feeding. J. ALQUER. *Ibid.* 390-435.—A review of the investigations carried out to date in the various labs, the results of which have already been reported to the Fr. Dept. of Agr. A. P.-C.

Fodder and quality of milk in meadow fertilization with calcium cyanamide. L. HERMANN. *Mischwirtschaft Forsch.* 9, 142-60(1929).—In a study on feeding fodder from land fertilized by Ca cyanamide the changes in acidity, density and fat content were observed. GEORGE R. GREENBANK

Determination of nitrite and sulfite in the presence of one another in meat products (SZABO) 7. Vitamins and palm oil in margarine (ERSTEIN) 27. Determination of Ca, Mg and P in feedstuffs (MORRIS, *et al.*) 7. Determination of cholesterol and lecithin. Evaluation of egg products (TILMANS, *et al.*) 11B. Treatment of canning wastes (etc.) (U. S. pat. 1,793,917) 27. Waterproof fibrous products for making milk bottles, etc. (U. S. pat. 1,793,839) 23. Apparatus and procedure for spray desiccation of milk (U. S. pat. 1,794,978) 1. Obtaining protein material from seeds (U. S. pat. 1,794,105) 27.

Fumigating flour mills with nicotine vapors. BRABENDER ELEKTROMASCHINEN

Gzs Brit. 337,311, March 1, 1929 A web of paper impregnated with nicotine is heated (as by passing it between elec heaters which are at a temp of 300-350°) without causing ignition but with carbonizing of the paper App is described.

Bread. STANDARD BRANDS, INC. Brit. 337,930, Oct. 11, 1929 A small quantity of urea (suitably with conjoint use of urease if desired, which enables a lesser proportion of urea to be employed) is mixed with the flour, yeast and other ingredients, and corn sugar, pure maltose and other sugars can be used in doughs contg urea, with production of a satisfactory crust color A small quantity of a slightly sol acid substance such as mucic acid (with or without an admixture of acid Ca phosphate) may be added to counteract the effect of liberated NH_3 on the pH of the mixt. The urea, urease and mucic acid may be preliminarily distributed in a filler such as flour, sugar or corn starch or soy-bean meal (which when used as a filler supplies urease)

Diastase baking powder. CARL H MEYER Swiss 141,861, Feb 4, 1929 Malt flour is mixed with powd. dextrose, preferably maize dextrose

Milk sterilizer. BERGEDORFER EISENWERK A-G ASTRA-WERKE Ger 517,217, July 11, 1929

Apparatus for sterilizing milk and similar treatments. BERGEDORFER EISENWERK A-G Brit. 338,422, Jan 9, 1929 Structural features.

Apparatus for pasteurizing milk, etc. ISAAC E COLVIN, GRANT N MANISON and HUGO G WALTERS (to Pfaunder Co) U S 1,794,670, March 3 Structural features

Apparatus for pasteurizing milk. CORNELIUS MOATENSEN U S 1,793,631, Feb 24 Structural features of an app with jacketed tubes

Apparatus for drying milk, etc. HANS LÜBER Ger 514,785, Jan 11, 1928 App with roller driers for milk and colloidal liquids is described

Thin brittle shell material for encasing ice cream, etc. WILLIS D. THOMPSON (one-third to Carl H Crawford) U S 1,794,741, March 3 A compn. which is brittle when chilled to below 22°, is formed of a suitably selected asphalt mixed with over 3 times as much paraffin and with a small proportion of stearin

Butter. E. G. N SALENTEUS and E S SALENTEUS Brit. 337,778, May 7, 1929 Whole milk or cream, during its passage from a centrifugal drum, is subjected to a vacuum of different degrees in different parts of the app (various details of which are described) Cf. C. A. 25, 546

Purifying butter, margarine and other fatty materials. R. BENDLIN and O. URBASCH Brit 338,053, Jan 9, 1929 Cleansing and working is effected by extrusion of the material into vermicelli-like threads which are immediately broken up by a jet of water, and the material is then passed to a settling vat and afterward is centrifuged The washing may be repeated and alk. or acid solns. used if necessary. The procedure may be applied to removal of excess of salt which has been added to butter as a preservative App. and various details of procedure are described.

Preserving meat by brine injections. G BEISSEK. Brit. 338,011, July 29, 1929 Various details of app. and procedure are described relating to injection of brines such as those contg NaCl, sugar and water in specified proportions into the vascular tracts of meat carcasses such as those of pigs

Meat-curing composition. ERNEST H WRIGHT. U. S. 1,792,561, Feb. 17. A meat-curing material is obtained by distg. hard wood at a temp of about 23-70° and discharging to the atm the gases released at these temps., and then gradually increasing the temp to about 480° and condensing the vapors evolved at these higher temps. App is described Cf C. A. 24, 2513.

Apparatus for sterilizing fruit juices. BERNHARD HUG Swiss 143,205, Nov. 18, 1929 Addn to 133,461 (C. A. 24, 443). The juices are heated by immersed electrodes.

Treatment of fruits and vegetables. F KNOX and C. WEST. Brit 337,422, July 31, 1929 For storage, transport or ripening, fruits or vegetables are placed in chambers of metal, glass or other materials provided with apertures covered with permeable material such as rubberized fabric of such character as to control the diffusion and dependent concn. of gases such as O, CO_2 and C_2H_4 in contact with the fruits or vegetables.

Preparing fresh fruit such as citrus fruits for market. ERNEST M. BROGDEN and MILES L. TROWBRIDGE (to Brodex Co) U. S. 1,795,275, March 3. For prevention of blue-mold decay, the fresh fruit is treated with a warm aq mold-inhibiting soln. of an alk. hydroxide (such as a 2% NaOH soln) having a degree of alkalinity nt least as high as that equiv. to 2 oz of Na_2CO_3 per gal, and sufficient amt. of the soln. is allowed to dry on the fruit to exert a decay preventing action Cf C. A. 24, 174.

Cellulosic coatings on articles such as fruits and vegetables. GEORGE W. BEADLE (to Cellacote Co). U. S. 1,794,751, March 3 A liquid hydrated cellulose prepn. such as cellulose xanthate soln is applied successively to different portions of the arti-

cle completely to cover it, and the article is then treated successively with a soln. of a neutral salt such as $(\text{NH}_4)_2\text{SO}_4$ and an acid soln. such as H_2SO_4 soln. and is washed and dried. App. is described. *Cl. C. A. 24, 2292*

Preventing blue mold of fruits such as citrus fruits. ERNEST M. BROGDEN and MILES L. TOWNBRIDGE (to Brogden Co.) U. S. 1,701,346, Feb. 24. In prep. fruit for the market is treated with a warm aq. mold-inhibiting soln. such as Na_2CO_3 of a degree of alkalinity at least as high as that of a soln. of Na_2CO_3 , 2 oz. per gal. of water. *Cl. C. A. 24, 901*

Apparatus (with a plurality of superposed conduits) for roasting coffee, cereals, etc. WM. L. PHILLIPS and JAMES R. NEAL (to Maxwell House Products Co.) U. S. 1,793,004, Feb. 17. Structural features.

Purine bases from cacao, etc. C. F. BORNHAEGER & SOHNE. Brit. 338,057, Jan. 7, 1930. Natural materials such as cacao are extr. with liquid SO_2 (suitably in a rotatable pressure vessel at a temp. of about 45°), and by filtering and cooling an ext. as thus prepd. a soln. of theobromine and caffeine and a layer of sepd. fat are obtained. The extn. may be repeated several times (preferably at higher temp., such as $100-105^\circ$). Tea dust may be similarly extr. for obtaining caffeine and theophylline, and mention is also made of the extr. of ground cacao shells. The material may be preliminarily treated with an alk. soln. *Cl. C. A. 25, 775*

Irradiating chocolate with ultra-violet. ERNST OTTENGRIFF. Swiss 141,862, April 9, 1929

13—GENERAL INDUSTRIAL CHEMISTRY

HARLAN S. MINER

Technology and industrial research. O. MANTZ. *Caoutchouc & gutta-percha* 28, 15175-6(1941).—A general discussion is given of the importance of research as evidenced by various cases which are cited.

Our foreign trade in chemicals in 1930. OTTO WILSON. *Ind. Eng. Chem.* 23, 430-4(1931)

American Society for Testing Materials, Standards. 1930, two parts, 2214 pp.—Standard specifications are given for open hearth C-steel rails; manuf. of open hearth steel girder rails; splice bars of various types of C-steel, track bolts and spikes of various kinds of steel; steel screw spikes and tie plates; structural steel of various types, and uses; rivet steel for boilers and for ships; boiler and firebox steel; steel plates of structural and of flange quality for forge welding; billet-steel and rail-steel concrete reinforcement bars; cold-drawn steel wire for concrete reinforcement; com. hot rolled bar steels; com. cold finished bar steels and shafting; C-steel bars for railway springs with and without special S₁ requirements; C-steel bars for vehicle and general purpose springs; silico-Mn steel and chrome-V-steel bars for railway springs; helical springs and elliptical springs for railways; C-steel and alloy steel forgings and blooms; billets and slabs for forgings, quenched and tempered C-steel and alloy-steel forgings for locomotives and cars; C-steel car and tender axles, cold rolled steel axles, wrought solid C-steel wheels for railways; steel tires and castings; C-steel castings for railroads; lap-welded and seamless steel and lap-welded Fe boiler tubes; welded and seamless steel pipe; C, high speed, and alloy tool steel; C-steel castings for valves, flanges and fittings for high temp. service; alloy-steel bolting material for high temp. service; forged or rolled steel pipe flanges and lap-welded and seamless steel pipe for high temp. service; Zn coatings on structural steel shapes, plates and bars and their products; Zn-coated (galvanized) sheets, telephone and telegraph line wire, tie wires, fencing; barb wire and steel wire strand; welded wrought Fe pipe; staybolt, engine-bolt and extra refined wrought Fe bars; hollow rolled staybolt Fe; common Fe bars; wrought Fe plates; wrought Fe rolled or forged blooms and forgings for locomotives and cars; Fe and steel chain; foundry pig Fe; cast Fe pipe and special castings; soil pipe and fittings; locomotive cylinders and wheels; malleable castings; gray-Fe castings; the arbitration test bar and tension test specimen for cast-Fe; W powder, spiegeleisen, ferro-Mn, ferro-Si, ferro-Cr, ferro-V, wire bars, cakes, slabs, billets, ingots and ingot bars of lake Cu and of electrolytic Cu; slab Zn; rolled Zn; pig Pb, Ni, Al for use in the manuf. of Fe and steel; Al ingots for remelting; Al sheet, phosphor Sn, phosphor Cu; silicon Cu; hot rolled Cu rods for wire drawing; Cu wire of various kinds; bare Cu cable; bronze trolley wire; an alloy of Cu 88, Sn 10 and Zn 2%; sand castings of an alloy of Cu 88, Sn 8 and Zn 4%; bronze bearing metal; bronze castings; compn. brass or ounce-metal sand castings; yellow brass sand castings; Mn bronze

Tentative specifications are given for: steel track spikes, open hearth fire plates, steel marine boiler plates, cold rolled strip steel, heat-treated C-steel helical springs; C-steel locomotive forgings, elec. cast steel stud link anchor chain, austenitic Mn steel castings, black and hot dipped Zn coated welded and seamless steel pipe, Zn coated chain link galvanized fence fabric, chilled tread cast Fe wheels, fire refined Cu other than lake, Al alloy (duralumin) sheet, Al Mn alloy sheet, Al base casting alloys in ingot form, Al base and Mg base alloy castings, sand castings of an alloy of Cu 80, Sn 10 and Pb 10%, Cu base alloys in ingot form, seamless Cu tubing, tests for compressive strength of portland cement mortars, high early strength portland cement, sand for use in lime plaster, gypsum breathing board, concrete aggregates, building brick and pipe, fire tests of building construction and materials, boiled linseed oil, tung oil, soybean oil, petroleum spirits, glazier's putty, dry-leached and orange shellacs, Al linseed and Al linseed powders, toxic ingredients in anti fouling paints, and nitrocellulose, synthetic AmOAc, AmOH and Et lactate, EtOAc, BuOAc, BuOH, Et propionate cum sizes of broken stone, broken slag, sand and gravel for highway construction, broken stone for waterbound and bituminous concrete bases, CaCl₂ for dust prevention, asphalt cement of various penetrations and for various uses, high- and low C tar cements for use cold in repair work, asphalt filler for brick pavements, mineral filler for sheet asphalt and bituminous concrete pavements, timber piles, wooden lures of various constructions, asphalt for roof coverings, rubber insulating tape, insulated wire and cable, specifications and test methods for asbestos tape for elec. purposes, rubber lined cotton fire hose, steam hose, tolerances and test methods for rayon and fur knit goods, tests for Cuban (jute) raw sugar lugs, chaser tire fabrics, thermometers for higher viscometers. *Tentative methods* are given for chem. analysis of Al, light Al alloys, rubber products and metallic materials for elec. heating, accelerated life test for metallic materials for elec. heating, test for thermoelec. power, flexure tests of concrete, testing brick for compression, flexure and absorption, tests for resistance of fire clay brick to thermal spalling action, tests for alk. or acidity, for leaching, for hygroscopic moisture and for oil absorption of pigments, sampling and testing shellac, lacquer solvents and diluents, detn. of Me₂CO ext. in dry lampblack and bone black, test for detn. of polishing lubricant in Al powder for paints; sampling petroleum and petroleum products, tests for color of lubricating oils and of petroleum by means of the Hylon Colorimeter, test for color of refined petroleum oil by means of the Saybolt Chromometer, tests for silt of crankcase oil, for distn. of crude petroleum, for exsorbable oil and moisture in paraffin waxes, for neutralization no. of petroleum products and lubricants, for penetration of greases and petroleum, for pytn. no. of lubricating oils, and for vapor pressure of natural gasoline, test for gr. of petroleum and its products by means of the hydrometer, test for S in motor fuels, naphthas and illuminating oils, tests for abrasion of gravel, for consistency of portland cement concrete, for ductility of bituminous materials and for residue of specified penetration; testing bituminous emulsions, test for size of anthracite; test for the detn. of the sp. gr. 34°/15.6°, of creosote fractions, test for coarse particles in bituminous materials by means of filtration; testing asphalt roll roofing surfaced with talc or mineral granules, likewise asphalt shingles surfaced with mineral granules, analysis of roofing felt for fiber compr.; compression testing of natural building stone; test for absorption and apparent sp. gr. of natural building stone, test for water absorption of slate, flexure testing of natural building stone and of slate; test for detg. the insulating qualities of slate, testing insulating varnishes, cable splicing and pothead compds., insulating paper, laminated sheet insulating materials, and varnished cloths and tapes; testing sheet and tape insulating materials for dielec. strength; testing elec. insulating materials for power factor and dielec. const. at frequencies of 100-1500 kilocycles; test for comparing the thermal conductivities of solid elec. insulating materials, testing insulating materials for resistance to impact, test for resistivity of insulating materials, test for hardness of soft rubber; testing grease wool, etc., for secured content, identification of textile fibers and their detn. in mixed goods; tension testing and compression testing of metallic materials. *Tentative recommended practice* is given for: bituminous paving plant inspection, thermal analysis of steel. *Tentative definitions* are given for: terms relating to magnetic testing, to the gypsum industry, to refractories, to petroleum, to timber preservatives, to textile materials and to methods of testing; the terms aggregate, coke and screen (sieve). *Tentative rules* are given for inspection of concrete and reinforced concrete work. *Tentative revisions* are given of standard specifications for steel track spikes, structural steel for ships, alloy-steel bolting material for high-temp. service, high-test gray-Fe castings, chilled cast-Fe wheels, ironore bearing metal in ingot form, various types of condenser

tubes and ferrule stocks, portland-cement hydrated lime for structural purposes, cement-concrete sewer pipe, broken slag for waterbound base and wearing course, shovel run or crusher run broken slag for waterbound base, block for various types of granite pavement, rubber pump valves, and tolerance and test methods for cotton fabrics. of standard methods of test for magnetic properties of Fe and steel, chem. analysis of Mn bronze, test for wetting point of fire-clay brick, making and storing specimens of concrete in the field, testing woven textile fabrics, sampling and analysis of creosote oil, test for distn. of creosote oil, testing elec. insulating oils and testing rubber products, of standard definitions of the term sand, of terms relating to the gypsum industry, refractories, paint specifications, coal and coke, textile materials and methods of testing. G. G.

Standards and specifications for nonmetallic minerals and their products. JOHN Q. CANNON JR. AND A. S. McALLISTER. Bureau of Standards, *Monographs* Pat. No. 110, (88) pp (1930).—Standards and specifications are given for coal and coke, coal-tar products and bituminous materials, charcoal, petroleum and its products, fuel and illuminating oils, lubricating and insulating oils and greases; asphalt and other bituminous materials, stone and stone manufactures, sand, gravel and slag, sand-lime brick, gypsum and other plasters, magnesite, cement and concrete, lime, construction work, flat glass, glass containers, table glassware, lighting globes and shades, lenses, chem. glassware, mirrors, clay, china and porcelain ware, earthenware and stoneware, bricks and tiles, natural and artificial abrasives, asbestos, chalk, mica, rare minerals, precious stones and imitations, S, magnesite, salt; graphite, fluorapatite and orthoclase feldspar silica. G. G.

Specifications and materials control for aircraft construction. H. A. BACHTA. *Proc. Am. Soc. Testing Materials* 33, Pt. II, 10-16 (1930). E. H.

The relative merits of gas, oil and electricity for industrial purposes. P. HOSKINSON. *Gas World* 94, Ind. Gas Suppl. 3 (1931).—The subject is discussed from the viewpoint of the furnace engineer. F. H. BERGMAN

Wooden barrels in the chemical industry. H. LOUIS F. HORN. *Chem. Markets* 23, 285-8 (1931), cf. C. A. 25, 1003. E. H.

Research in drying. M. LUKY. *Arch. Wurmwart.* 12, 21-3 (1931).—L. gives very brief descriptions of the drying investigation of various German drier constructors, and of the Moscow drying lab. ERNEST W. THIEL

The process of evaporation. A. V. BLOM. *Forsbr-Zig* 36, 573-5 (1931).—The phys. laws governing evapn. are discussed. For single and mixed solvents, the rates of evapn. are expressed by the equations, $y = at^n$ and $y = \int_0^t f(t) dt$ resp., in which y = amt. of liquid evapd. in time t , a , the original amt. of liquid, n , f and t , constants which depend upon the conditions of evapn. The value of n is independent of the presence or absence of non volatile matter. G. G. SWARD

Industrial poisoning with aniline and similar substances. ANGELO BONANNO. *Deut. Z. ges. gerichtl. Med.* 16, 242-5 (1931).—Cases are described of poisoning with aniline, nitro- and dinitrobenzene, nitrochloro- and dinitrochlorobenzene, phenylhydrazine and phenetidine and benzene. FRANCES KRASNOW

Poisoning with dilute tin chloride solution. R. ZETVICK AND H. WIELSCH. *Deut. Z. ges. gerichtl. Med.* 16, 194-9 (1931). FRANCES KRASNOW

Carbon monoxide poisoning. H. HUG. *Deut. Z. ges. gerichtl. Med.* 16, 72-88 (1931).—Case report. Discussion. FRANCES KRASNOW

Observations of carbon dioxide poisonings in the disaster at Neurode. G. W. PARADE. *Deut. med. Wochenschr.* 55, 1283-8 (1930).—Clinical observations were made after the mine disaster at Neurode. ARTHUR GROLLMAN

Comparison of heat-insulating materials for industrial furnaces and boilers. VIKTOR PASCHEN. *Arch. Wurmwart.* 12, 9-11 (1931).—If k is the heat cond. of any substance, and p its price per unit vol., then for an infinite flat plate the most economical material (other things equal) is that which has the lowest value of k/p . For actual solid bodies, especially furnaces of small dimensions, a higher power of k is required. In continuous use, k/p does well, in most cases k^2/p is still better as a discriminating quantity. ERNEST W. THIEL

Rosin in cable-impregnating compounds. J. P. MILLWOOD. *Proc. Am. Soc. Testing Materials* 30, Pt. II, 82-91 (1930).—Saturants used to impregnate the paper windings around the Cu conductor of a cable have usually contained rosin or rosin oil or both, dissolved in petroleum oils or greases. The rosin serves to improve consistency and tack of the compd., secures more complete and lasting impregnation and better retention of the saturant, prevents "bleeding" when the ends are cut, prevents move-

ment of the saturant within the cable, and doubtless insulation troubles. The use of rosin, however, serves to limit the operating voltage and current loading of the cable owing to the low dielectric strength and high dielectric loss of rosin, causing the cable to become almost hot when heavily loaded. When a wax seed and blended with the proper oils, however, this effect is minimized. The use of rosin has been found to result in the formation of small, rather than large voids within the saturant. Small voids are preferable, because gas contained in them is less apt to bubble than gas in larger voids. Properties of which rosin, to be used in cables, should possess are: (1) dryness, (2) absence of oxidized rosin, (3) freedom from dirt, impurities and abnormal free rosin, (4) constancy of the above and other properties. H. K. SATZGER.

Paragutta, a new insulating material for submarine cables (KREYER) 30. Final report of I. M. M. representatives on the above and a subcommittee of the British Engineering Standards Association (KREYER, SCHWAB) 2. Thermal properties of compressed liquid water (KREYER) 2. Some final values for the properties of saturated and superheated water (KREYER, SCHWAB) 2. Some additional volume data for superheated steam (SCHWAB, KREYER) 2. The higher alcohols of the paraffin series (as emulsifying agents) (SCHWAB) 10. Treating bluish materials with rubber, etc., for making insulation, etc. (BULL. pat. 337,700) 28. Apparatus for "neutralizing" poisonous gases (U. S. pat. 1,700,913) 1.

Annual Reports of the Society of Chemical Industry on the Progress of Applied Chemistry, 1928. Vol. XV. London: Society of Chemical Industry. 623 pp. 12s. 6d. Reviewed in *Chem. Trade J.* 59, 201 (1931).

Rossett, Henry N. The Chemical Technology of Steam-Raising Plant. London: Edward Arnold and Co. 210 pp. 12s. 6d.

Chemical Manufacturers' Directory of England, Wales and Scotland, 1931. London: Simpkin Marshall, Ltd. 201 pp. 4s. 6d. Reviewed in *Chem. Trade J.* 59, 202 (1931).

Reuss, Kurt. Kolloidtechnik. Halle (Saale): W. Knapp. 212 pp.

Reuss, Kurt G., and Cox, William J. Hypothesis and Mechanics of Materials. New York: J. Wiley & Sons, Inc. 352 pp. \$4.50.

Reynolds, Robert H. Chemistry of Engineering Materials. 2d ed. New York: McGraw-Hill Book Co., Inc. 681 pp. \$4.50.

Rizzo, Pietro. Trattato completo di geometria gubules. Material, empirico-facile e filosofico. Madrid: Agencia Espanola de Libreria. 201 pp. Ptas. 8.50.

Vittoria, G. Vittoria, et al. Dictionari di microscopia e di chimica applicata. III. Metallurgia. 2d ed., revised and enlarged. Milan: Pirelli Hoepli. 1. 60. Reviewed in *Chem. Trade J.* 59, 201 (1931); cf. *C. & E.* 21, 2314.

Separating Liquids. A. G. FOR BREITENHAUSEN. Subs. 142,835. Oct. 20, 1928. Akin to 174,531 (C. & E. 24, 1654). App. is described for sepr. liquids lighter than water, e. g., CCl₄, from waste water.

Filtrating alkaline liquids. ALUMINUM INDUSTRY A. G. Bull. 337,659, Jan. 17, 1929. See Oct. 202,131 (C. & E. 24, 4870).

Clarifying suspensions. LEONARD FRANK. Ger. 517,055, Feb. 22, 1928. The settling of suspensions is accelerated by treating them in turn with a plant ext. such as carageen moss and a salt of Ca or Sr, particularly CaSO₄. The treatment of resins is related to.

Distillation of volatile substances. I. G. FARBENFABRIK A. G. (Hans Traubsen, inventor). Ger. 514,743, Feb. 13, 1928. In the sepn. of volatile substances from non-volatile volatile substances by distn. with a liquid whose b. p. lies below the distn. temp., part of the liquid is added, while the other part forms a heat-insulating layer between the liquid being distilled and the liquid about to be added. App. is described. The method is useful for wining oils and fats.

Solvents for cellulose esters and ethers, resins, fats, waxes and rubber. DEUTSCHE HYDROKARBON A. G. Ger. 517,097, Nov. 27, 1928. Esters of cyclopentanone carboxylic acid with monohydric alcohols are used, alone or with other solvents, etc. The esters may be prepared by treating esters of sulpho acid with NaOH.

Mixing materials for cement, pigment, fertilizer, acid phosphate or dye manufacture, etc. J. H. MORROW (to Fuller Co.). Bull. 339,123, Oct. 3, 1928. In prep. materials, as for cement manuf., blending is effected by passing the material from a mill radially to a plurality of "silos" in bins (a thin layer being successively added to the contents

of each of these) and by suitable withdrawal and redistribution of the material. Various details of app. arrangement are described.

Projecting hydrolyzable fluids such as titanium chloride into contact with air. JOHN T. REMY. U. S. 1,702,029, Feb. 17. The fluid is discharged into the air (by a device which is described) simultaneously with a surrounding sheath of another non-reactive fluid such as dry air, CO_2 , or CCl_4 in order to avoid clogging of the exit nozzle.

Gas filters for breathing purposes. GRASMAN K. E. H. STAMPE (to Otto H. Dräger). U. S. 1,702,030, Feb. 17. Filters for removing NiH_2 and H_2S are formed with Ca_2Cl_2 on a solid carrier substance such as activated charcoal treated with Pb nitrate soln.

Absorbent for ammonia in refrigerating apparatus. RALPH M. BUFFINGTON (to Frigidaire Corp.). U. S. 1,702,628, Feb. 17. SrCl_2 and NH_4NO_3 are used together. Cf. C. A. 25, 1920.

Condenser (suitable for refrigerating apparatus). ROLLIN M. HYDE (to McCord Radiator & Mfg. Co.). U. S. 1,704,002, March 3.

Forming sheet insulating material such as wall board comprising a filter cake and associated deposit. FRANK W. YOUNG. U. S. 1,704,433, March 3. Materials such as fibrous pulps may be used and app. and various details of manuf. are described.

14—WATER, SEWAGE AND SANITATION

EDWARD BARTOW

Sterilization of drinking waters at Saigon and at Cholou. J. GUILLERM. *Arch. Inst. Pasteur Indochine* 1929, No. 10, 36-55. *Chimie & Industrie* 25, 9(1931).—A description of the water supply its treatment and control. The water, which is obtained from an underground sheet which is imperfectly protected from contamination by an overlying layer of clay of varying thickness and insufficient impermeability, is rendered perfectly safe by chlorination with bleaching powder at a cost of 0.0008 franc per m^3 . Sterilization is controlled by counts of the aerobic and anaerobic bacteria, test for (and count of, if present) indole-forming *B. coli* and identification of other species of bacteria and protozoa.

St. Louis has model water system. L. A. DAY. *Water Works Eng.* 83, 731-2, 819-24(1930), cf. C. A. 24, 5899.—This well-illustrated article with diagrams and statistics gives detailed descriptions of the equipment, treatment, operation and maintenance of the Chain of Rocks and the Howard Bend purification plants and their systems of distribution.

New York to rebuild century-old Croton Aqueduct as a safety measure. WM. W. BRUSH. *Water Works Eng.* 83, 736-8, 831(1930).—A brief history of the old Croton Aqueduct is given and its present condition discussed. The daily demand on the municipal water system within the city is 900 m. g. d., while at times outside communities need 30 m. g. d. The yearly growth of demand will average 30 m. g. d. The present sources of supply are the Catskill system delivering 600 m. g. d., the New Croton Aqueduct, 280 m. g. d., and local sources of Long Island and Richmond, 120 m. g. d. To avoid a possible water shortage, since a new large source of supply will not become available for at least 6 or 7 years, the weak sections of the old Croton Aqueduct will be rebuilt so that its present delivery of 40 to 45 m. g. d. can be doubled. The probable cost will be between 1 and 2 million dollars and the work will be finished in 1931.

Toronto constructing \$14,000,000 tunnel. EARLE W. GAGE. *Water Works Eng.* 83, 654(1930).—The new filtered water tunnel, across the city, will follow the shore-line of Lake Ontario. It will be 7 ft. in diam. for more than 5 miles and 6 ft. in diam. for nearly $4\frac{1}{2}$ miles. The first section will deliver 90 m. g. d. into the heart of the city and 50 m. g. d. will be received at its western end. In addn. 60 m. g. d. will go directly into the city mains from high lift pumps at Victoria Park. A pure water reservoir here will hold 12,000,000 gal. and a new service reservoir in the northern part of the city will contain 50,000,000 gal.

Observations of water supplies of London and Paris. EDWARD BARTOW. *J. Am. Water Works Assoc.* 23, 267-71(1931).—In one of the plants visited excess lime treatment was used and the plant operation included lime recovery and removal of tastes and odors by activated C. In others compressed air for filter washing and ozone for purification were considered novelties.

Unusual construction problems met in two Far Eastern water works. M. E. BARNES. *Water Works Eng.* 83, 625-6, 658-61(1930).—Engineering difficulties in the

Far East are discussed. Every building in Bangkok, located 40 km. from the sea on the Menam Chao Phya River, is built on mud. This river carries enormous loads of silt and directly receives the sewage of Bangkok via its many canals and the pollution of many thousands who live in house-boats. Its chief objection as a source of water supply is that the water is made brackish from the flow of sea water past Bangkok upstream in the dry season. The water supply is taken from a natural reservoir created out of a branch stream 41 km. above Bangkok. A canal, poiced to prevent diversion and pollution, carries the water to a modern water works erected on the outskirts of Bangkok. This plant contains sedimentation tanks, Jewell filters and equipment for distribution at an equalized pressure. It has proved its value in cholera outbreaks. By a suitably placed dam, a large catchment area located in that part of the Federated Malay States adjacent to the Island of Singapore was drawn upon for impounding purposes.

C. H. BADGER

Increasing supply of mountain city uncovers some unique problems. GEORGE D. NORCOM. *Water Works Eng.* 83, 629-30(1930).—The water consumption of Bluefield, W. Va., has averaged about 1.2 m. g. d. Because of an 18 months' drought in 1925-6 the supply became inadequate. Since then the storage water capacity has been increased an addnl. 127 m. g. by construction of a dam and reservoir at the mouth of Horton Hollow. This reservoir now receives also the Gravity Springs supply which formerly was usually improperly chlorinated because of its remoteness and fed directly into the mains. The reservoir supply now flows by gravity to the Ada filtration plant, which also treats water obtained from a 45-m. g. earth reservoir supplied by Ada Spring. The treatment consists of aeration, coagulation, sedimentation, rapid sand filtration and chlorination. Ada Spring is unusual because of its cure and flow, which has varied from 0.175 m. g. d. to over 30 m. g. d. Provision was made against leakage through the soft limestone rock on which the new Horton dam was built by cleaning and grouting all weak spots in the rock and forcing cement grout into deep holes drilled into the rock to fill seams and crevasses. The Beaver-Pond Bailey supply, which also was formerly chlorinated and pumped directly into the mains, has been retained as an auxiliary source in case of interruption to the Ada-Horton development. Daily chem. and bacteriol. tests are made for the control of purification.

C. H. BADGER

Water differences—geographically speaking. C. C. HUMPHREYS. *Southeast Water Works J.* 12, No. 11, 15(1931).—H. discusses unusual water conditions and turbidity, acid, Fe and phenol at Pittsburg, Pa., and algae at Ithaca. At Tampa, Fla., 300 p. p. m. color is removed by SO_2 and Cl_2 .

O. M. SMITH

How the planets affect our water supply. CALLED M. SAVILLE. *Water Works Eng.* 83, 733, 847-8(1930).—A general discussion is given of the long-period studies of climatology and meteorology. There is a greater tendency for rain in the quarter after full moon, especially with the moon in perigee, than with it in apogee. While the correlation between tree growth, as shown by tree rings, e. g., the giant Sequoia, and climatic cycles seems to be definitely established, that between sun spot phenomena and terrestrial climate has not as yet been accepted.

C. H. BADGER

Waterworks intakes and the screening of water. JOHN W. CUNNINGHAM. *J. Am. Water Works Assoc.* 23, 258-66(1931).—The paper applies principally to flowing streams. For effective screening the design and arrangement of the intake are of great importance. Racks—intake screens—(2 to 4 inches per in.) and fine screens (8 meshes per in. or finer) are considered with reference to their best selection and location. Some reference is made to cleaning.

D. K. FRENCH

Water purification for color removal. A. S. BEHRMAN, R. H. KEAN and H. GUSTAFSON. *Paper Trade J.* 92, No. 8, 121-3(1931).—Modern methods of color removal are briefly reviewed, and results are given of an investigation into the nature of the coloring matter in water and new methods of removal based thereon. It was found that: (1) The particles of coloring matter are negatively charged. (2) The particles of coloring matter are of a colloidal nature, though some of them may approach mol. dimensions and consequently be in true soln. (3) The coloring matter in surface waters is essentially org. in nature. (4) In many cases the color deepens considerably with increase in pH of the water, particularly when passing from acid to alk. conditions. (5) The color is destroyed to a considerable, but variable, extent by chlorination. The excess Cl_2 must be removed after destruction of the color in order to prevent corrosion; and with all the waters tested it was found that dechlorination by filtration through granular hydrocarbons removed not only the Cl_2 , but also the whole of the residual color. Filtration through hydrocarbons, without previous chlorination, removed all or nearly all the color; but owing to the necessity of reinvigoration, it is usually more economical to remove as much color as possible by coagulation and filtration under carefully controlled

pu condiments followed by chlorination and then dechlorination if necessary. An attempt is being made to find a method of revivification that will be so simple as to make it practical and economical to dechlorinate the water directly by treating with hydrocaro, without preliminary coagulation, filtration or chlorination. A. P. C.

Treatment of water for ice manufacture. DANA BUKES, JR. Univ. Ill. Eng. Expt. Sta., Bull. No. 219, 114 pp (1930).—This work done in cooperation with the Illinois Research Commission embraced a study of methods for producing clear ice from natural waters high in dissolved salts, particularly Na_2CO_3 and NaHCO_3 . With neutralization of the water with alum or H_2SO_4 to convert the carbonates to the less troublesome sulfates and with the use of a specially designed ice can providing ample and effective air agitation, a marketable ice fairly free from opaque zones was produced. If, however, the core water reaches a critical concn. of salts it must be replaced with a new supply. A second line of attack involved the utilization of the exchange properties of Mg zeolite to replace Na salts in the water with the equiv. Mg compds. Subsequent liming, followed by alum or acid treatment, produced from NaHCO_3 water supplies one from which satisfactory ice could be made. H. L. OLIV.

Galvanized iron pipes for city water supply. J. DE GRAAFF. *Verlag Mededeel. betreffende Volksgezondheid*, July 1930, 8 pp.—The requirements for galvanized iron piping are discussed. It should be nonporous. A test for porosity of the Zn coating is to pour a 10% gelatin soln. with 1% $\text{K}_4\text{Fe}(\text{CN})_6$ on a part of the inside of the pipe which has been cut in half. After 24 hrs. at room temp. blue spots in the solidified gelatin indicate breaks in the coating. The thickness of the Zn layer should be sufficient. It is detd. by dissolving the Zn in HCl ($d = 1.2$) to which 20 g. SnO_2 per l. has been added. This acid does not attack Fe. A simple app. for the detn. is described: a 2- to 3-cm. Cu tube, 3 cm. long, is filed off to fit the inside of half of the galvanized pipe. Proper sealing is obtained by pouring asphalt around the Cu pipe. Acid is poured in through a funnel in one hole of a double bore stopper closing the Cu pipe, H₂ evolved escapes through the other hole and is collected and measured in a bottle. A min. of 600 g. Zn per sq. m. of galvanized surface (one side) is required. The Pb content of the Zn used should be less than 0.5%. On the nature of the water depends whether the Zn will be effective in reducing corrosion. If the water causes a CaCO_3 deposit after some time, the corrosion chances will be diminished. If no CaCO_3 is deposited, the Zn will dissolve in time and corrosion take its course. From an example is shown that in water of low hardness as much as 3 g. Zn per l. is present after standing 1 night in galvanized piping. Water left for 30 hrs. in a new pipe even had 7 mg. per l. Zn gives at best a temporary protection. B. J. C. VAN DER HOFVEN.

Subsoil management, particularly from the viewpoint of water mains. BRABANT, MOYAERTS, et al. *L'eau* 24, No. 1, 5-8 (1931).

Why water mains break. FRANK A. McINNIS. *Water Works Eng.* 83, 529-30, 550 (1930).—Rigid bearing, or uneven settlement of the supporting ground, electrolysis, earthquake, frost, effect of salt marsh land and abnormal chem. character of the pipe itself are some causes given for the fracture of cast iron pipes. C. H. BADGER.

Iodine content of Lettish waters in relation to the improvement of crops. J. KUPCIS. *Acta Univ. Lettenis Kim. Fakultat. Seriya 3*, 425-48, 445-53 (in German) (1930).—K. gives complete analyses for Cl samples of Lettish waters and discusses the sources of the I in the various samples. He discusses the need of I in growing crops and points out various causes for the decrease in the amt. of available I. C. J. WEST.

Detailed instructions for the performance of the dissolved oxygen and biochemical oxygen demand tests. EMERY J. THERIAULT. *U. S. Pub. Health Repts. Suppl.* No. 90, 44 pp (1931).

Method of determination of the chlorine demand of any water. D. K. PATILLO AND P. D. WEST. *Paper Mill* 54, No. 5, 4 (1931).—Cl demand is defined as the amt. of Cl that will combine chemically, or possibly biologically or bacteriologically, with any given water, white water or sewage to so change the state of oxidizable matter that it will not require addnl. O. The method consists essentially in titrating 500 cc. of the sample with hypochlorite soln. contg. 0.0001 g. available Cl per cc. until a blue color is obtained with iodized starch soln. used as outside indicator, adding a further slight excess of Cl and then detg. the residual Cl with a standard colorimetric Cl testing outfit. A. PAMNEAU-COUTURE.

Early use of chlorine. WM. BOBY. *J. Am. Water Works Assoc.* 23, 283-4 (1931).—A paper published earlier is discussed. The equipment used as well as the results obtained in the first Belgium installation in 1902 are described. The process employed was invented by Dr. Maurice Duyk, official chemist to the Belgian Govt. D. K. F.

Progress in superchlorination treatment for taste prevention at Toronto, Ontario.

N J HOWARD. *J. Am Water Works Assoc* 23, 387-95(1931)—See C A 24, 4878.

G. G

Preammoniation of the filtered water supply of Cleveland, Ohio. J W ELLMS. *J Am Water Works Assoc* 23, 400-7(1931)—See C A 24, 4878

G G

Eijkman test for the bacteriological examination of water. MARIO MAGALHAES. *Arch hyg Rio de Janeiro* 4, No 2(1930), *U S Pub Health Eng Abstracts* 11, W, 19(Feb 21, 1931)—Eijkman's method employs a glucose peptone medium, to which the water under examn is added in the ratio of 1 8 and incubation carried out at 46°. In 18-24 hrs a distinction can be made between *B coli* from cold- or warm blooded animals. In addn to being selective for *B coli*, it inhibits common water organisms and bears a correlation with the tests of indole production, the citrate and uric acid reaction, and may thus prevent contamination of water when *B coli* are of non human or cold-blooded origin

C R FELLERS

Dental deficiencies and drinking water. R W KERR. *J Am. Water Works Assoc* 23, 214-29(1931)—A dental defect, known generally as mottled enamel, is described, and the location of communities in which it is found is mapped. It is thought to be due in some way to the drinking water used by the victims. No explanation can be given at the present time

D K FRENCH

Lead poisoning from drinking water in Leipzig. KRUSE AND M FISCHER. *Deut med Wochschr* 56, 1814-8(1930)—A clinical account is given of a recent epidemic

ARTHUR GRÖLLMAN

Water supply of a steel plant. T J McLOUGHLIN. *Proc Eng Soc West Penn* 46, 295-305(1930)—Extremely large vols of water are necessary, one ton of steel requiring 14,000 to 17,000 gallons water. The main intake of a steel plant is one of the most vital spots. Extreme care is exercised to prevent injury by floating debris and ice. Modern variable speed traveling screens insure freedom from floating material. Much of the water in an iron region, especially the Pittsburgh district, periodically contains considerable free H_2SO_4 . A considerable portion of the water must be neutralized with milk of lime to prevent corrosion of the distribution system. Boiler-feed water usually is treated river water. A few of the smaller plants use a hot lime-soda process; some use a continuous cold lime-soda treatment; a few are zeolite plants, but by far the most use an intermittent lime-soda process followed by filtration. Internal treatment is rapidly becoming recognized as an economical and effective method of scale prevention

WAYNE L DENMAN

Softening a well-water supply. N. T. VEATCH, JR., AND B. L. ULRICH. *J. Am Water Works Assoc* 23, 272-5(1931)—The Manhattan, Kans., supply is softened by lime and alum. With chem softening alone a hardness of 463 p p m. was reduced to 165 p p m. With recarbonation in addn this was lowered to 124 p p m.

D. K. F.

Water supply for high-pressure boilers. CHAS R. HAZEN. *Pulp Paper Mag Can* 31, 195-8(1931)—The formation and effects of scale and the factors favoring corrosion are briefly reviewed, and the usual processes of water softening are outlined. They do not furnish a water suitable for the very exacting requirements of high-pressure boilers. The prime essential of such boilers is clean evap surfaces; this can be obtained by the properly controlled addn of Na_2HPO_4 , either in the boiler or before feeding, and with or without a preliminary softening treatment according to the nature of the raw water

A. PAPINEAU COUTURE

Boiler-scale prevention by the use of trisodium phosphate in modern boiler management. PAUL KOEFFEL. *Chem.-Ztg* 55, 58-9(1931)— Na_3PO_4 is the ideal softener for use in modern high pressure boilers. It is inexpensive and efficient and requires no special equipment to add the softener. Contact during only 1 hour at 70° is required to cause pptn of the Ca and Mg salts

C R FELLERS

Developments in the treatment of acid mine drainage. LEWIS V. CARPENTER AND ALFRED H. DAVIDSON. *Bull West Va Univ Sci Assoc* 2, No 4, 50-7(1930)—The acidity from coal mine drainage has created a serious problem in West Va. The amt. of drainage is approx 25% of the rainfall. Ordinarily the flow of mine drainage will be 1000 gal per day per acre of coal exhausted and is independent of the nature of the season. The acidity cannot be predicted, as often one mine will give an alk drainage, while another nearby will be highly acid. The conversion of S to acid is rapid and is hard to explain by the chem process of oxidation. Some hold that the action of anaerobic bacteria is responsible. Treatment has consisted of the addn of powd lime, limy marl, limestone and magnesite, but is not used to much extent practically. The Kaplan-Reger process is said to remove all chemicals in the water, forming a blue pigment, but little is known about it. Very little is known concerning the details for proper disposal by dilution. On a cost comparison basis, limestone is the cheapest. If allowance is

made for capitalized cost of the plant, drying of sludge and no return for sale of sludge, the total cost per 1000 gal is \$3.00. At this cost the process is not feasible. To be successful, the process must allow disposal of the by products. WAYNE L. DENMAN

Eliminating odors and nuisances about the sewage-disposal plant. F. R. HASSLER. *Southeast Water Works J.* 12, No. 10, 17 (1931).—The use at Perry, Okla., of NH_3 3 to 7 lbs per million gals together with an av. of 25 lbs of Cl_2 in domestic sewage preceding digestion in sep. sludge digestion tanks removed H_2S odor and produced a stable effluent. O. M. SMITH

Reclamation of treated sewage. R. F. GOURNEY. *J. Am. Water Works Assoc.* 23, 230-40 (1931).—By reclamation is meant treatment of such a high grade of efficiency that the water as well as the solids can represent credits in water supply economy. In Southern California this need of reclamation is of the utmost importance. While some work has already been done, much more is necessary before all its difficulties are overcome. D. K. FRENCH

Digestion of sewage solids at high temperature. THEO. F. DODDIS. *Munich Sanitation Z.* 121-4 (1931).—Samples of fresh sewage solids seeded with ripe sludge were digested at the thermophilic temp. of 55° and at room temp. Detns. of the nature and quantity of gas evolved, pH , dry solids, ash, volatile matter and no. of bacteria were made upon duplicate samples. Thirty days was the optimum digestion period. From 18% to 20% more gas was evolved from solids digested at thermophilic temps. than at room temp. CO_2 , O , CH_4 , and N were present. The CH_4 and CO_2 contents fluctuated inversely. Gas produced in either case was 65-75% combustible. A pH of 7.3 was the optimum for digestion at thermophilic temps. Total bacteria present and solid detns. show little difference under the 2 conditions. GUNNAR L. KESLO

Utilization of sewage gas. H. F. KETTEL. *Monats-Bull. Schweiz. Ver. Gas-Hausfach* 10, 293-8 (1930).—Sewage gas from settling tanks of the city of Zurich (700,000 m^3 annually) contained 66% of very pure CH_4 , (30% CO_2 , N_2 and a trace of H_2S). Possibilities for the utilization of CH_4 , chlorination, oxidation, conversion into fuel oil, etc., are considered in the light of recent patent literature. K. H. ENGEL

The mechanism of the activated-sludge process of sewage disposal. E. C. C. BALY. *J. Soc. Chem. Ind.* 50, 22-6T (1931).—The sign of the elec. charge on sewage colloids can be detd. by flocculation of suspensions of fine powders of known charge and by means of cataphoresis. Both methods indicate that sewage colloids behave as if they were electropositive when the pH is less than 8.3 and electronegative when it is over 8.3. Sewage colloids have an isoelec. point at about pH 4.6 which shifts toward the side of greater pH in the presence of electrolytes. The colloids in ordinary sewage with pH 7.4 are therefore electronegative. The bacteria of sewage carry an electronegative charge, and when their concn. is sufficient they are flocculated by electropositive colloids. When their concn. is insufficient they exist as bacteria and colloid complexes, when the colloids are electropositive. The success of the activated-sludge process is due to the much enhanced charge on the bacteria when in their state of max. activity. The necessity for this max. activity would disappear if the colloids were electropositive and operation of the process is proposed at pH 5.8-6.0. W. H. BOYCE

Advantage of sludge circulation and vacuum degasification as applied at the Cleburne plant. CLYDE C. HAYES AND CHESTER COHEN. *Proc. 12th Texas Water Works Short Course* 1930, 161-6. *U. S. Pub. Health Eng. Abstracts* 11, S. 20-1 (Feb. 14, 1931), cf. C. A. 24, 3301.—The Cleburne plant consists of a grit chamber, Imhoff tank with central gas vent, dosing chamber and Miller siphon and a sprinkling filter unit. Inadequate sludge digestion capacity, 0.5 cu. ft. per capita, foaming difficulties, odor troubles, filter bed clogging and inadequate purification of the effluent were some of the difficulties of operation formerly. Some of the theories advanced for the present hastened digestion with degasification and sludge circulation are: (1) That through the circulation of the digested sludge the end products which are inhibitory to bacterial life and are produced through decompn., are effectively removed in the form of gases distributed through the tank liquor. (2) thorough and continuous mixing of the ripe and fresh solids aids in speeding digestion and maintaining a desirable H_2 ion concn. and secures a thorough mixing of the bacterial flora of the sludge, (3) the entrained gases are more easily liberated through the stirring action and more quickly carried off through the vacuum maintained on the gas vents, (4) the reduction of the gas concn. in the top area of the gas vent through the use of the vacuum system decreases pressure of the gas through the entire body of the liquid, (5) the maintenance of the vacuum on the gas vents serves to further increase the efficiency of the Imhoff slot, and thus tends to prevent belching of solids upward into the sedimentation compartment, (6) the scum and foam which usually accumulate on the top of the gas vents can be mechanically drawn off, at

least partially digested, which saves addnl. labor cost incidental to this process; (7) in spite of the reduced sludge-digestion capacity the tanks are producing a fairly well digested sludge which can be discharged on drying beds, without any fear of odor nuisance or fly breeding. While economic use of captured gas is not considered, as the primary function of the plant is sewage purification, a 4-hp gas engine can be successfully operated from the sewage gas. Complete details of plant operation are given with description of plant units. Conclusions. The vacuum system of degasification and sludge circulation demonstrated the following advantages: (1) Continuous circulation of the digesting sludge is advantageous in hastening the process of org. decompn. (2) the circulation of the sludge through the use of the vacuum pump is most economical and satisfactory. (3) the putrefactive odor from the Imhoff tank is entirely controllable, (4) foaming difficulties are avoided, (5) labor cost for scum removal and gas vent control is practically eliminated, (6) the reduced size of tank made possible through this vacuum system and sludge circulation will decrease the cost of Imhoff tank construction, greatly simplify and reduce operating expenses and minimize legal complaints. C. R. FELLERS

The role of protozoa in activated sludge. ROBERT CRAMER. *Ind. Eng. Chem.* 23, 309-13(1931), & *C. A.* 24, 2819.—Clarification of sewage can be accomplished by aeration if aerobic bacterial life, live protozoa and oxygen in soln. are present. The biochem. nature of the process is indicated. Material decompd. by aerobic and anaerobic bacterial life in sewage would remain in a colloidal, non-settleable state if the protozoa did not resynthesize it. F. W. TANNER

Lagooning sludge. EDGAR WHEEDER. *Proc. 12th Texas Water-Works Short School* 1930, 153-6, *U. S. Pub. Health Eng. Abstracts* 11, 8, 18(Feb. 14, 1931).—The sewage-disposal plant at Dallas, Texas, lagoons the sludge drawn from the digestion chambers of the Imhoff tanks as well as scum from gas vents and skimmings from the surface of the flow chambers or settling basin. During extended rainy weather the sludge beds were filled and the digestion compartments became overloaded, and, when accompanied by foaming, became large septic tanks. The difficulty was remedied by running the excess sludge into a shallow slough or lagoon nearby. No odors came from the lagoon. The yield of wet sludge per million gals. was 7.6-8.0. C. R. FELLERS

Determination of carbon in sewage and industrial wastes. F. W. MOHLMAN AND G. P. EDWARDS. *Ind. Eng. Chem., Anal. Ed.* 3, 119-23(1931).—The C detn. of Friedemann and Kendall (*C. A.* 23, 2733) has been applied to domestic sewage and trade waste and can be completed in 3 hrs. Results with carbohydrates, soap and amino compds. indicate satisfactory accuracy for the purpose. When org. C/total N on raw sewage, corn products, stockyards and tannery wastes is 3.98-42.7, org. C/biochem. O demand is 0.55-0.65. This indicates a relation between org. C and 5 day O_2 demand and supports the theory that the first stage of the O_2 demand curve is a C oxidation. An activated-sludge effluent is in a different stage of oxidation and cannot be directly compared, the corresponding org. C/biochem. O demand being 4.1. IOSTER DER SNEEL

Fundamentals of public health law. JAMES E. BAUMAN. *U. S. Pub. Health Repts.* 46, 631-41(1931). J. A. KENNEDY

Emergency sanitary measures following a flood disaster. PAUL S. FOX. *Am. J. Pub. Health* 21, 291-6(1931).—F. details his experiences following the floods which occurred on the Rio Grande in New Mexico in Aug. and Sept., 1929. J. A. K.

Amelioration of atmospheric pollution. HOWARD W. GREEN. *Am. J. Pub. Health* 21, 237-41(1931). J. A. KENNEDY

Legal aspects of water pollution. LEO T. PARKER. *Munic. Sanitation* 2, 129-31(1931). C. L. KILSO

Precautions needed in the ammonia-chlorine treatment of swimming pools. A preliminary study. LYNN L. JENSEN AND HENRY R. WELSPER. *Ind. Eng. Chem.* 23, 32-4(1931).—The use of chloramine for sterilization makes possible a high residual Cl content without the unpleasant taste, odor or irritating effects of free Cl. With the recirculation of swimming pool water as generally practiced, the NH_3 content is cumulative and reduction to NO_2^- occurs in certain cases. These products interfere with the color test for available Cl. Use of an antichlor will destroy chloramine. Excess color formed under such conditions should be ignored when considering Cl values for sterilization. The apparent fallacy of this treatment should not outweigh its advantages. LOWARD S. HOPKINS

Creoline as a larvicide. DURNY. *Ann. soc. belge mtd. trop.* 8, 23-5(1928); *U. S. Pub. Health Eng. Abstracts* 11, 14, 2(Feb. 14, 1931).—Lab. expts. made with heavy tar oil showed 200 p. p. m. was the min. effective dosage to prevent mosquito breeding.

The oil must be applied every 10 days. It is of particular value in shallow pools rather than in deep water. C. R. FELLERS

Progress in the treatment and disposal of human wastes. R. E. McDONNELL. *Munic. Sanitation* 2, 114-8 (1931)

Treating of canning wastes. ANON. N. Y. State Dept. Health, Div. Sanitation, *Bull.* (unnumbered) 1930, 100 pp; *U. S. Pub. Health Eng. Abstracts* 11, 5, 19-20 (Feb. 14, 1931).—Analytical data are tabulated in 28 tables. Suggested treatment plant designs are given. Conclusions: Screening through wire screens having 20 to 40 meshes to the in. is a necessary preliminary treatment in all cases and adequate treatment when the final diln. is greater than 1 to 100. Forty sq. ft. of effective screen area is necessary for a 2 line cannery discharging 100,000 gal. of wastes per working day of 12 hrs. About 80 cu. ft. of screenings are produced per 100,000 gal. of wastes treated. Trickling filters are very effective in the treatment of screened canning wastes when the rate of filtration is less than 0.5 million gal. per day throughout the working day, and this treatment is adequate when a final diln. is greater than 1 to 5. Rates of filtration of 2 million gal. per day throughout the working day are sufficient when the final diln. is greater than 1 to 20, provided serious pooling is prevented through the periodic application of lime to the applied wastes to prevent the prolific growth of fungi on the filter. Chem. pptn. of canning wastes with 7-10 lb. of lime and 3-5 lb. of FeSO_4 per 1000 gal. of wastes is sufficient treatment when the final diln. is greater than 1 to 25. Although the cost of constructing chem. pptn. equipment may be less than that of trickling filters, nevertheless the cost and difficulty of operation, cost of chemicals and the less satisfactory effluent justify the use of trickling filters in most cases. Final chlorination of filtered or coagulated and pptd. canning wastes modifies the org. matter and is equiv. to the removal of about 30% of the residual unstable org. matter. Final chlorination, therefore, permits of the discharge of filtered canning wastes into streams with final dilns. as low as 1 to 3, provided the rates of filtration of 0.5 million gal. per day are used. Final chlorination of filtered or coagulated canning wastes also materially reduces or entirely suppresses the fungus growths in the streams receiving the effluent, thus eliminating an objectionable result of pollution. The required dose varies from 8 to 20 p. p. m. Cl_2 depending upon the degree of preliminary treatment of the wastes to assure a residual of at least 0.5 p. p. m. in the effluent. Every effort should be made to reduce the water consumption of canneries and to prevent the wastage of the product so as to reduce the strength and quantity of wastes to be disposed of. The relatively high cost of treating these wastes, which are 5-15 times as costly as raw domestic sewage, makes it very desirable that efforts be made to recover by-products from the wastes and thereby reduce the net cost of their disposal. C. R. FELLERS

Treatment of tannery wastes (ANON) 29. Corrosion of steel water pipes by stray electric currents (ROTHE) 9. Titrimetric determination of small quantities of NH_3 . With particular attention to water analysis (HAGEN) 7. Supplying artificially ionized O for ventilation (U. S. pat. 1,793,799, 4).

KLUT, HARTWIG. Untersuchung des Wassers an Ort und Stelle. 6th ed. Berlin: J. Springer 180 pp. M. 7.80

Apparatus suitable for filtering water through sand, etc. HAROLD D. ELFRETH (to Cochrane Corp.) U. S. 1,794,841, March 3. Structural features.

Softening water with base-exchange material. GEORGE BORROWMAN. U. S. 1,793,670, Feb. 24. The water is percolated through a granular bed comprising lignite or brown coal contg. an exchangeable alkali metal until the latter is "more or less exhausted," and the bed is then revived by percolating a NaCl soln. through it and washing it, and is then further used for water purification.

Base-exchange compounds. GENERAL ZEOLITE CO. Brit. 337,768, July 27, 1929. Reaction is effected between an alk. metal silicate, an alk. compd. of an amphoteric hydronide and a non-alk. compd. of an amphoteric hydroxide, to form a gel including substantially the whole reaction mixt., e. g., a soln. of Al sulfate contg. added H_2SO_4 is mixed with a Na silicate soln. and a soln. of Na aluminate is added. An acid or caustic alkali or both may be added to the reaction mixt. and the use of Na zincate and other materials is also mentioned. The gel may be dried and washed. Cf. C. A. 24, 909.

Water-softening apparatus. WALTER H. GREEN (to International Filter Co.). U. S. 1,794,765, March 3. Various structural details are described of an app. operating with a layer of zeolitic material superposed on a layer of non zeolitic material such as gravel.

Domestic water vessels containing base-exchange material for softening water placed in them. E. BARNHART and ETABLISSEMENTS PERMATS & PAPEL. Brit. 338,021, Dec. 2, 1920. Various modifications are described of an app. of the general character of that described in Brit. 233,004 (C. A. 23, 1978).

Apparatus for purifying boiler feed water by sedimentation, etc. EDWARD J. BROOK. U. S. 1,794,353-6, March 3. Various structural details are described.

Preventing furting. JEAN SCHNEIDER. Swiss 141,897, May 24, 1930. Furring is prevented in water heaters by addn. of H_2CrO_4 and a colloid. In the example, an amount of H_2CrO_4 and colloidal SiO_2 is added. Cf. C. A. 25, 750.

Preventing furting in boilers. FARR HADGECOCK. Ger. 514,815, Dec. 3, 1928. Colloids are added to the water which is then made to deposit its fur, by heating under pressure, before admission into the boilers.

Distributor box for sewage-disposal system. JOHN F. HARRIS. U. S. 1,758,123, March 3.

Sending tank for removing sludge from sewage. GUSTAV R. ROEY (to Chalmers Belt Co.). U. S. 1,758,143, March 3.

Apparatus for conditioning air with water sprays, etc. A. T. LEWIS, H. P. GANT and R. R. TAMMARENO. Brit. 338,018, Nov. 28, 1922. Structural features.

Treating waste liquors to remove sulfide sulfur. COCHRANES, LTD., R. S. JONES and S. FRANKSON. Brit. 337,533, Sept. 23, 1922. The liquor is first treated with acid until neutral or very slightly acid, in the presence of the SO_3 radical, as by the use together of $ZnSO_4$ or $MgSO_4$, $NaHSO_4$ and H_2SO_4 , and on mixing the materials part of the S settles out. The remaining liquor is then treated with caustic alkali in the presence of a Zn or Mg salt until it is alk. to phenolphthalein, when $Zn(OH)_2$ or $Mg(OH)_2$ and the rest of the S settle out.

Recovering oil from waste water. FRYMANT'S CASTILLAN. Swiss 141,439, Sept. 4, 1922. App. for the sepn. of $C_{12}H_{22}$ oil and petroleum from waste water, including a revolving drum with internal compartments, is described.

15—SOILS, FERTILIZERS AND AGRICULTURAL POISONS

J. J. FIEDORER AND M. S. ANDERSON

Summary of results of German soil analyses up to 1920. LUTER. Z. Pflanzenernähr., Pflanzeng. u. Bodenk. 53, 337-9, (1930).—The degree of soil acidity and the Fe_2O_3 and K_2O contents are given in tabular form.

The temperature correction in the hydrometer method of mechanical analysis of soils. CHARLES ROBERT. Soil Science 31, 83-92, (1931).—A study is made of factors affecting the mech. analyses of soils by use of a hydrometer. Soil suspensions of a particular soil at different concns. but at the same temp. will not give identical results when the temp.-correction factor of 0.55 is applied at all concns. Concordant results are obtained only when the correction applied is considerably smaller at the lower range of the hydrometer than at the higher range. A sliding-scale correction table is suggested.

M. S. ANDERSON

Variation of the colloidal material extracted from the soils of the Miami, Chester and Cecil series. R. S. HARRIS and GLEN EMMERTON. U. S. Dept. Agr., Tech. Bull. 229, 1-25, (1930).—A specially designed agitator is described in which the soil suspensions were prepd. A 75% wgt. of the colloidal material from these suspensions was obtained by means of a supercentrifuge. The colloids derived from the various samples of each series show a marked constancy, but they show an equally marked divergence between the different series. The largest amt. of colloidal matter is found in the B horizon. The H am. concn. is on the soil side in all but 1 horizon. With few exceptions the org. matter, S and P are most abundant in the A horizon. SiO_2 increases with depth in most of the Chester profiles, whereas the reverse is true for the Miami and Cecil series. The active base constituents vary in the Miami profiles in the reverse order to their variations in the Chester and Cecil profiles. The Fe_2O_3 content is higher in the lower horizons of all 3 profiles. The relative constancy in compn. of the colloids of each series is noted and discussed. The colloids of the different soil series are, in general, very different. If the series are arranged in the order of the SiO_2 : R_2O_3 ratio the order becomes Miami, Leonardtown, Chester and Cecil. This order also generally holds for the 4 major constituents if the reciprocal character of the relation is kept in mind. Combined H_2O varies inversely with the SiO_2 : R_2O_3 ratio. The univalent and bivalent bases do not follow this ratio very closely.

J. R. ADAMS

Interaction between ammonia and soils, as a new method of characterizing soil colloids. AMAR NATH PURI *Soil Science* 31, 63-7(1931)—Soils react very slowly with bases. Equal conditions can be reached quickly only when a large excess of alkali is added and the excess removed by leaching or by pptn. as some insol. compn. Completely unsatd. soils will combine with an amt. of NH_3 , which may be considered equiv. to the base-exchange capacity. The method used for NH_3 adsorption consists of placing the H-satd. soil in a flat dish in a vacuum desiccator over $\text{N} \cdot \text{NH}_4\text{OH}$ for 2 days. The dish is then removed and placed over 90% H_2SO_4 in vacuum for 48 hrs. The NH_3 retained is detd. by distn. with lime. The pH values of soils contg. CaCO_3 are about the same after dil. acid treatment and subsequent NH_3 satn. as for the soil in the natural condition. M. S. ANDERSON

Investigations on the organic matter of soils. Their determination and their importance as nitrogen reserves. VINCENT, HERVIAUX AND GAUDIN *Ann. agron.* 47, 654-71(1930)—Humic materials extd. from a soil vary with the chem. nature of the alk. solvents, their concn., the temp. and the reaction time. The concn. of the alk. soln. used affects the concn. of the humus. High alk. concns. break down humic material and too weak solns. produce variable results. The humic materials and humus reduce 0.1% K permanganate solns. in a fairly const. manner. The combinations of N and their probable role in soil fertility are summarized. The necessity of stating precisely in soil analyses, the existing humic and N combinations is emphasized. The detection and detn. of humic materials are essential. The humate solns. obtained were analyzed for total and sol. humic materials by oxidation by K permanganate. It is wished to study the ammonification of sol. org. materials, the extn. in the cold for a protracted time with a weak alk. soln. is preferable. It is advised that 20 g. of finely ground, decalcified soil be agitated mechanically for 2 hrs. with 500 cc. of 1% soda soln. followed by 24 hrs. standing with occasional shaking by hand. P. W. MARSH

Some methods for detecting differences in soil organic matter. EDWARD C. SNORREY *U. S. Dept. Agr. Tech. Bull.* 211, 1-25(1930)—The dark color of the alk. exts. of soils is not due to the absorption of O during the process of extn., and such color cannot be considered an indication of the presence of phenolic compds. Differences in the shades of color of humus exts. are proof of differences in the kind of org. matter, but a comparison of the depth of color of the exts. cannot be used as a measure of the quantity of org. matter present. The N in the various fractions of an alk. ext. of a soil and the quantity of O absorbed when air is passed through the ext. vary with the soil. The presence of 2 classes of org. nitrogenous compds. in soils is indicated. These are chitin and its deriv., glucosamine, and some member of the indole group. The results of acid hydrolysis of soils from a N standpoint and the use of alc., water and alternate alc. and water exts. are discussed as means of showing differences in soil org. matter. The formation of asoporphine when some soils are heated with PhNH_2 is described, and this formation is suggested as a means of showing the presence or absence of some as yet unknown org. soil constituents. A short discussion of a no. of general reactions that may be applied to soil exts. or fractions, such as the pyrrole and fluorescein reactions, are presented. W. H. ROSS

The replaceable cations in the soil and the plant. E. K. GEDROIZ. *Udobrenie i Urozhai (Fertilizers and Crops)* 2, 464-75(1930)—A chernozem soil was satd. with the cations, Mg, Ba, Mn, Co, Ni, Cu and H. It was used in pot expts.—700 g. per pot—on mustard with a complete fertilizer and without. No crop was obtained with any one of the enumerated cations. In another series of expts., with the same soils, 10 g. CaCO_3 was added, and the exptl. plant was oats. No crop was obtained from Ba, Ni and Co soils. A small crop was obtained with the Cu soil (0.39 g. of dry matter against 5.4 g. on the regular chernozem soil with CaCO_3), a slightly better crop—1.56 g.—from the Mg soil, and 1.6 g. from the Mn soil. The H soil gave a normal crop. The next series of expts. was conducted with solid satd. with Mg, Ca, Sr, Cd, Fe^{++} and Fe^{+++} , Al and H. All pots received a complete fertilizer and one series received also an addn. of 10 g. CaCO_3 . Without CaCO_3 , no crop was obtained from the Cd and Fe^{+++} soils, some small yield was obtained from the Mg, Al and H soils. The Sr soil gave just as good a crop as the soil satd. with Ca, which in turn gave just as good a crop as the chernozem. Apparently Sr can replace Ca in the nutrition of plants. The pots which also received CaCO_3 gave a crop with all the cations, except Cd. The crops on the Fe^{++} and Fe^{+++} soils were small, and on the Al and Mg soils they comprised 50% of the normal. Thus the Ca cation is one of the most important exchange cations from the standpoint of plant nutrition. From earlier expts. G. cites the effect of a K-satd. soil. Without Ca the K was not available. On H-satd. soils only CaCO_3 brought about a normal crop, addns. of CaSO_4 had no effect. This is due to the acid reaction produced by the

formation of H_2SO_4 . It was also found that part of the Ca may easily be substituted by Mg without any ill effects. J. S. JOFFE

Determination of the percentage base saturation of soils and its value in different soils at definite pH values. W. H. PIERRE AND G. D. SCARSETT. *Soil Science* 31, 99-114 (1931).—A comparison is made of the base satn. of soils when detd. by the $Ba(OH)_2-NH_4Cl$ method and by the $Ba(OAc)_2-NH_4Cl$ method. The $Ba(OH)_2-NH_4Cl$ method gives appreciably higher total exchange capacities than does $Ba(OAc)_2$. Soils that have been limed show much smaller differences by the 2 methods. Soils of various origins at like pH values may show considerable differences in percentage satn. In general, highly weathered soils such as those from the Piedmont Plateau and of the Coastal Plains have a lower degree of satn. than do those from the Coastal Plain Black Belt and the Glacial and Loessial Province. There appears to be no relation between org. matter content of soils or the nature of the bases present in the exchange complex and the percentage base satn. at like pH values. The avidity or strength of acids present varies with different soils. In general, highly weathered soils have weaker acids than the less weathered ones. There is a good correlation between the avidity of soil acids and the percentage of base satn. of soils at pH 4.80. M. S. ANDERSON

The buffer capacity of soils, methods of determining it and its practical value. P. KUCHINSKY. *Ann. versn. inst. Akad. Lening. Gory-G-N* 9, 77-104 (1929).—K. detd. by the method of Jensen the buffer capacity of the loess like soil in the vicinity of Gorki. The effect of microrelief on the buffer capacity and the variations in the profile were studied. At the same time the exchangeable bases were detd. by the Kappen method as well as by the hydrolytic acidity. The illuviation horizons had the highest buffer capacity, the eluviation horizons the lowest. When the clay (particles <0.01 mm.) was removed, the buffer capacity was practically zero. The humus portion of the soil is high in buffer capacity. With the increase in buffer capacity the base-exchange capacity is increased. The higher the buffer capacity the higher the fertility of the soil. On the slopes the buffer capacity is low in the depressions it is high. There is a correlation between buffer capacity and hydrolytic acidity. J. S. JOFFE

Soil acidity and its practical significance. I. DE V. MALHERBE. *S. Africa J. Sci.* 27, 253-69 (1930) (in S. African Dutch).—The causes and effects of soil acidity are reviewed. Exchange acidity is detd. according to Daikuhara (*C. A.* 9, 500) with methyl red (or phenolphthalein) as indicator (cf. also Gov. *et al.*, *C. A.* 24, 874), the method consists in shaking 100 g. soil with 250 cc. N KCl soln. for 1 hr., filtering and titrating with 0.1 N NaOH. Comparative results of this acidity with pH are given for several soils. The hydrolytic acidity (Kapper, Bodenanalytik, Berlin 1929) is detd. by titration after shaking with $Ca(OAc)_2$. It is always higher than the exchange acidity. A table for some 150 S. African soils of different natures gives % rock (>0.04 mm.), % clay + humus, pH and both acidity values. Of the total, 32% is in the neutral class (pH 5.99) and does not need lime. Thirty-five % is weakly acid (5.50 to 5.99) and needs slight lime addn. if sandy. The exchange acidity is a good indication for this need, it is higher for the humus soil than for the clay soil because of buffer effects. The same holds for the stronger acid soil groups (19% of 5.00 to 5.49 pH). As a rule lime is always required for soils with more than 2 cc. acidity. Special cases of fertilization are discussed. The type of crop grown is a decisive factor in the matter of liming, the most acid-sensitive crops are corn, barley, sugar beet, cabbage, beans, peas and lucerne. The amt. of lime required for improvement of the soil is discussed and methods for calcn. are given. Many references are included. B. J. C. VAN DER HOEVEN

Active soil acidity and crop yields. I. T. SOLDATOV. *Ann. versn. inst. Akad. Lening. Gory-G-N* 9, 141-65 (1929).—From a series of field expts. in various rotations the relation between the pH of the soil and its crop response has been established; for rye $5.94 = 0.014$, for oats $6.05 = 0.030$ and for potatoes $5.92 = 0.13$. J. S. JOFFE

The potassium thiocyanate method for determining soil acidity. YUTAKA KAMOSHITA. *J. Imp. Agr. Expt. Sta. (Japan)* 1, 55-8 (1930); cf. *C. A.* 16, 158.—Treatment of soil with an alc. soln. of KCNS produces a red color if the soil is more acid than pH 6.88, and the depth of color is proportional to the acidity of the soil. At the same pH values, the color increases with exchangeable acidity, since this is accompanied by increase in Al_2O_3 and Fe_2O_3 and decrease in CaO and MgO brought into exchange. The exchangeable Fe^{+++} of soils is therefore considered responsible for a part of the exchange acidity. K. KITSUTA

Change in soil acidity. E. PASCH. *Z. Pflanzenernähr. Düngung u. Bodenk.* 9B, 309-26 (1930).—On limed soils green manuring lowered the pH about 0.6-1.0. Physiologically acid and alk. reacting fertilizers increase acidity less. Barn manure did not produce increased acidity. H. B. SEXTON

Colorimetric determination of pH values in alkaline soils. P. KAMRAN. *J. S. African Chem Inst* 13, 59-63(1930).—Colorimetric methods for detg pH values gave unsatisfactory results. Discrepancies are probably due to hydrolytic effects brought about by the increased water soil ratio necessitated by the prepn of a filtered soil ext. B C A

Production of soil carbon dioxide. TH. OASLER. *Z. Pflanzenernähr. Düngung u. Bodenk* 9B, 415-21(1930).—Forced ventilation of soil in pots produced larger amts of CO_2 than was in soils not so treated. Increased CO_2 production resulted in better plant growth. Aeration of soils by tile drainage may result in greater production of CO_2 , through the decompn of soil org matter and thus increase crop yields. H B S

Temperature variation of cultivated soils. OTTO HAUSER. *Z. Pflanzenernähr. Düngung u. Bodenk* 9B, 252-67(1930).—Soil temp at various depths has an important influence on plant growth. The change in soil temp at the surface, at depths of 5, 17.5 and 30 cm with respect to air temp from January to July, is reported. During July the temp of the surface and 5 cm depth was slightly higher than the av. air temp. At 17.5 cm depth the temp is nearly equal to the av. daily temp, except when rapid changes in air temp take place. At 30 cm depth the temp is a little lower. Variation in daily soil temp is also caused by rain, liming of soil, cultivation and covering with straw and paper. H B SIEM

Organic compounds associated with base-exchange reactions in soils. W. T. MCGROGGE. *Ariz Agr Expt. Sta., Tech Bull* 31, 215-51(1931), cf. *C. A.* 25, 161.—The arid soils of the Southwest contain only small quantities of org matter. The compn and reactions of this org matter were investigated. Lignin and lignin like substances are the most abundant and active base-exchange agents in arid soils. The exchange capacity of the lignin present in soils is not a const quantity but varies in different soils. The same is true for ligno-humates, although the range of variation is not so great. The aq. alkali sol ligno-humate has a much higher exchange capacity than the alc. alkali sol lignin. Leaching the lignin and ligno humate with H_2O increases the exchange capacity, probably by hydrolysis. The quantity of lignin extd from soils by alc. alkali represents a comparatively small percentage of that actually present. Max soly may be obtained by digesting with aq. alkali at increased pressures. Titration of lignic acid (II-satd lignin) and ligno-humic acid (II-satd ligno-humate) with KOH and Ba(OH)₂ indicates that the lignin mol is dibasic, and ligno humate tetrabasic. The absorption of the base of an acetate by lignin is equal to that required as hydroxide to neutralize the II satd salt. Ionization of the acid and basic salts of lignin and ligno-humates was detd by measuring the cond. of their resp. solns at several dilns. The Na and K org complexes show rather high ionization, while those of Ca, Ba and H are very low. The effect of a common ion on ionization and base exchange in org matter was studied. The influence of a common ion upon replacement by another base is appreciable, but, except for Ca, it is less outstanding than in inorg zeolites. Xylan exhibits, to a slight degree, the property of base replacement, but this is not of sufficient magnitude to account for the greater exchange capacity of the ligno-humate as compared with lignin. Synthetic humus, prepd from xylan or cellulose, like that prepd from sucrose, yielded materials with rather high base-exchange capacities. Xylan, lignin and ligno-humate absorbed color from basic fuchsin soln, and this color was replaceable by the base of a neutral salt soln. Green manure, from ground dry alfalfa, shows an appreciable base-exchange capacity, a large part of which is not easily destroyed by digestion with H_2O_2 . The base-exchange capacity of ground alfalfa was increased 4 fold by spontaneous decompn. Just as with soils, the extn. of ground alfalfa with alc or aq. NaOH yields a lignin of high base-exchange capacity. The lignin content of org matter and highly org soils is a linear function of the base-exchange capacity, while there is no relationship between hemicellulose or cellulose and the exchange capacity. The exchange capacity of the org fraction increases as the org matter passes through successive stages of decompn. in the soil. A bibliography is appended. C R FALLERS

The exchangeable bases in Malmesbury slate soils. M H SLABBER. *S African J Sci* 27, 270-9(1930) (in S Afr. Dutch).—Loam soil samples from the Malmesbury and Stellenbosch districts were studied, they are of slate origin. Top soil (9" deep) and the next layer (18" deep) were used. The Kelley method was used for detn of exchangeable base (Kelley and Brown, *C. A.* 20, 2232). The results are tabulated for Ca, Mg, K and Na and include pH . The av. is 58 mg equiv total exchangeable base for hill soils, it is 35 mg equiv for lower grounds. These low figures are due more to the colloidal condition of the soils than to an unsatd condition, the low humus content is significant. The relative ratios of the bases are discussed, in the low grounds Na and

K predominate. Even if more K is available for plant food, the condition of soil with high univalent ion content is generally unfavorable, the clay fraction flocculated. The exchangeable base content of the soils studied is only an av. of the total base present in them, a figure lower than customary. B. J. C. v.

Behavior of water in a drained field. JON ROTUR. *Z. Pflanzenernähr. u. Bodenk* 9B, 512 (1930).—In certain drained soils 26-29% of the total free water was found in the drainage water. H. B.

Further studies on the relationship between the concentration of the soil and the physicochemical properties of the leaf-tissue fluids of cotton. J. HARRIS AND TRUMAN A. PARSONS. *J. Agr. Research* 41, 767-88 (1930).—The general a pos. correlation between the salinity of the soil and the osmotic pressure measured in terms of l. p. depression, specific conduct., Cl content and sulfate content of the leaf tissue fluids of both Pima Egyptian and Lone Star upland cotton. The relation between the elec. resistance of the soil and the concn. of all solutes of the leaf fluids, as measured by l. p. depression is higher in both the upland and Egyptian varieties than the correlation between the salinity of the soil and the cond. of the fluids. In general the correlation between the analytically detd. concn. of the solutes studied in the soils and the total solutes of the tissue fluids, measured in terms of conductance, is higher than that between these ions and the various measures of elec. content of the tissue fluids. In the upland variety the correlations between resistance and total solutes of the tissue fluids are higher than those between resistance and the chlorides of the tissue fluids while the reverse is true with the Egyptian variety. In the Egyptian variety the correlations between soil resistance and solutes are higher than those between soil resistance and tissue fluid sulfate. This is not as marked and may be reversed in the upland type. J. R. v.

The determination of assimilable nitrogen in soils by the growth of grasses. H. D. WAGNER. *Bull. Arch. Landw. Abt. A. Pflanz.* 5, 163-200 (1930).—The N method for detg. P and K is extended to the detn. of assimilable N in soils and for Westerwold Raygrass was found to give the best index of the N content of the soil to utilize the N supplied in any of the common fertilizers. W. GORDON.

Soil nitrates as a guide to the nitrogen needs of vegetable crops. H. D. OHIO Agr. Expt. Sta., *Bimonthly Bull.* 149, 65-8 (Mar.-Apr. 1931).—The nitrate plate test of Morgan (*C. A.* 24, 3972) was used to advantage in testing N deficiency in soils cropped to various vegetables. For good crop growth the soil should contain more p. p. m. of nitrates. Fresh strawy manure usually tested 6-7 p. p. m. of N. A concn. of 60 p. p. m. did not injure spinach. N is often not liberated from rich greenhouse soils fast enough for tomato and cucumber production, the crop suffering from N hunger. The remedy is the application of NaNO_3 . C. R. v.

Determination of plant-food content in soils by the citric acid and other methods. J. HANSEN AND R. BALKS. *Z. Pflanzenernähr. Düngung u. Bodenk* 9, 73 (1930).—The citric acid method for detg. the available P_2O_5 and K_2O of soil is practically as good as the biological methods. No method is without objection. chem. detns. have the advantage that they can be carried out quickly. Each, however, gives an indication in regard to the deficiency, medium content and abundance of phosphate and potash of the soils. H. B.

Permeability of iodine in some economic plants. R. C. MALIPUTRA. *Proc.* 12, 1-22 (1931).—Soil was analyzed for I. Portions were then treated with amounts of that element and used for growing a variety of plants. Carrots take most, tomatoes the least. There is no relation between the amt. stored by the plant and the amt. added to the soil. Root crops take up more than leaf crops. Iodine is localized chiefly in the roots. A soil pH of 6 is the optimum. Seventy-two refs. are appended. H. A. v.

The nutrition of seedlings and the effect on root formation of cereals. K. Z. *Pflanzenernähr. Düngung u. Bodenk* 9B, 433-45 (1930).—During the germination of seed, enzymes change carbohydrates, fats and proteins into water-soluble compounds. After enzymic splitting of the large molecules, oxidative enzymes then change the products into smaller molecules with the loss of CO_2 . If the moisture content of the soil is insufficient for germination, enzyme activity may still proceed with the evolution of small amounts of CO_2 , thus decreasing the total solids in the seed. Under these conditions resultant seedling may be greatly weakened. Growth of seedlings may be promoted if the seed is moistened (25-40% H_2O content) from 2 to 3 days in a saturated atmosphere. To prevent the action of oxidative enzymes, the seed is best stored in a vacuum. A seed is best moistened by a fine spray and not by dipping it in water to prevent

of water sol substances Seed thus treated gives a better root system and a larger crop yield H B SIEMS

Crop rotation. P MONTEA *Z. Pflanzenernähr Düngung u Boden* 9B, 529-52 (1930)—The total plant food and the ratio N P_2O_5 K_2O , representing the plant food removed by different field crops, were not const for the same plant, but depended on weather conditions, tillage, shading soil conditions and insect injury Beets removed the most plant food, nearly twice as much as rye and barley Potatoes and oats followed beets in amt of removal Wheat extd less K_2O Leguminous plants removed practically as much P_2O_5 and K_2O as potatoes and oats H B SIEMS

A mathematical study of the decrease of crop yields. J DUDLEY GREAVES *Soil Science* 31, 115-22(1931)—Further study of an equation developed for the study of crop yields tends toward the establishment of the 2 assumptions on which the formula was based The rate of increase of crop yield with increase of the deficient element is proportional to the magnitude of the deficiency of the limiting nutrient from an optimum concn The time rate of depletion of a deficient element, provided none is added from an outside source, is proportional to the product of the soils content of the deficient element and crop yield M S ANDERSON

Chemical and microbiological study of Lufkin fine sandy loam in relation to productiveness. E B RYNOLDS *Texas Agr Expt Sta., Bull.* 421, 5-30(1931)—In field expts the nitrifying power of the soil was correlated positively and significantly with the yields of cotton and corn The nitrifying capacity was a better index of soil productivity than any other method studied The nitrifying power of the soil was also positively correlated with the total N, total P_2O_5 and available P_2O_5 of the soil The nitrifying capacity of the soil in the lab was not affected by the season Continuous cropping of the soil to corn lowered the nitrifying capacity The application of nitrogenous materials, cottonseed meal and manure, and also P_2O_5 as superphosphate or rock phosphate, increased both nitrate production and nitrifying capacity Liming greatly increased nitrate production C R FELLERS

Influence of various non-nitrogenous compounds on the growth of certain bacteria in soils of low productivity. H J. CONN AND MARY A DARROW *N Y. Agr Expt Sta., Tech Bull.* 172, 3-40(1930)—Two soils of low fertility were sterilized and inoculated with *B. globiforme* Conn and with 2 related species This organism has N requirements similar to those of green plants, and it was believed that the test might show why the N in these unproductive soils was not available to plants The bacteria did not grow in the soils unless NH_4 salts, nitrates or other forms of readily available N were added The addn to the soils of the hydroxides, sulfates, carbonates or phosphates of the strong alkali metals also made the N of the soil available for the bacteria It is probable that the cause of the unavailability of the N is adsorption by soil colloids In this case the bacteria acted like green plants C R FELLERS

The influence of inorganic iodine compounds on the metabolism of some important soil bacteria. HELMUT KAHNAN. *Zentr. Bakl. Parasitenk., Abt.* II, 82, 494-518 (1931)—As much as 1 g per l of KI, KIO_3 , NaI or $NaIO_3$ decreased the production of CO_2 Concns of 1 salts below that gave slight increases of CO_2 The respiration rate was slightly decreased, iodate having a greater effect than iodide 1 salt concns between 1 and 10 g per l inhibited denitrification Below that there was little effect The results for N fixation and urea destruction were similar. JOHN T MYERS

Presence of *Azotobacter* and absence of *Thiobacillus thiooxidans* in peat soils. C E SKINNER AND I J NYGAARD *Ecology* 11, 558-61(1930)—*Azotobacter* is active in peat soils only at $pH > 5.9$ *A. chroococcum* was found in the more basic soils and *A. Beijerinckii* in those approaching the acid limit *Thiobacillus thiooxidans* was not found in peat except where S had previously been used as a fertilizer B C A.

The determination of the potassium needs of the soil by *Aspergillus niger*. H NIKLAS AND H POSCHENRIEDER, with the collaboration of G VILSMEIER. *Wiss Arch Landw., Abt. A, Pflanze* 5, 152-65(1930); cf *C A.* 24, 5915—Ten soils of different types were investigated as to their P_2O_5 and K_2O content by the Neubauer method, and by the *Aspergillus niger* method, and the yields of various crops grown on these soils were detd (1) without fertilization, (2) with fertilizer deficient in (a) N, (b) H_2PO_4 and (c) K, and (3) with complete fertilization The results obtained from the *Aspergillus niger* method are to be published later with the crop yields reported in this article. W GORDON ROSE

Potash content of Pfalz soils and conclusions in regard to fertilizer recommendations. M KLEIN AND O ENGELS *Z. Pflanzenernähr Düngung u Boden* 9B, 409-16 (1930)—The root sol. K_2O in various soils was detd by the Neubauer method Tentative "limiting values" of mg of K_2O per 100 g dry soil are given as follows for grain 30,

vegetables and clover 30, grapes 50, tobacco 50, pastures and meadow grasses 30. In many cases the crop yield could be increased by adding potash to the soil, although the soil contained more than the "limiting value" of potash. H. B. SIEMS

Nature of the yield curve [of plants] with increasing applications of potash. If WIESSMANN AND K. NAUMANN *Landw. Jahrb.* 72, 105-23 (1930).—The yield curve of oats and rye grass with increasing supplies of potash approximated to the logarithmic type of Mitscherlich (cf. C. A. 24, 224). Later sowing of the seed caused slight increases in the "effect factor" of potash. The mean values of the "effect factor" of potash with rye grass and with oats were similar, but with the former crop there was a slightly decreased value with the higher applications of potash. In general values of the "effect factors" for potash were slightly smaller than those of Mitscherlich. B. C. A.

Utilization of increasing amounts of potash on the crop yield of several varieties of barley. H. LIESEGANG *Z. Pflanzenernähr. Düngung u. Bodenk.* 9B, 397-409 (1930).—Four varieties of barley grown in pots varied in their ability to utilize increasing amounts of potash as shown by crop yields, although the amt. absorbed under the same conditions was practically equal. H. B. SIEMS

The effect of sodium nitrate and ammonium sulfate on the availability of soil potash and soil phosphoric acid. M. GRACANIN *Z. Pflanzenernähr. Düngung u. Bodenk.* 9B, 300-9 (1930).—The amt. of phosphate and potash absorbed by plants from the soil as measured by the Neubauer seedling method is increased noticeably when $(\text{NH}_4)_2\text{SO}_4$ is added to the soil. On extremely acid or alk. soils this effect was not noticeable and in some cases was neg. NaNO_3 caused an increase in absorption of P_2O_5 by plants from all soils except one with a pH 8.35. The increase in the absorption of potash did not parallel the P_2O_5 absorption in all cases. H. B. SIEMS

Remarks concerning A. Nemec's treatise "The evaluation of the results of soil analyses with respect to the need of phosphate fertilization." M. V. WRANGELL. *Wiss. Arch. Landw., Abt. A, Pflanze* 4, 635-40 (1930).—Polemical. Cf. N. (C. A. 25, 1935) and *Deut. Landw. Presse* 57, No. 23 (1930). W. GORDON ROSS

The determination of soluble soil phosphates. B. DIRKS AND F. SCHEFFER. *Wiss. Arch. Landw., Abt. A, Pflanze* 4, 641-51 (1930).—A discussion is given of the methods available for the prepn. of a soln. contg. the easily sol. soil phosphates and of the factors that must be considered in detg. the amt. of phosphate that must be supplied for various crops. W. GORDON ROSS

The mechanical analysis of finely divided natural phosphates. LYLE T. ALEXANDER AND K. D. JACOB. U. S. Dept. Agr. *Tech. Bull.* 212, 1-24 (1930).—The pipet method for the mech. analysis of soils (C. A. 24, 1452) was successfully applied to the analysis of finely divided natural phosphates. For the analysis of phosphates, pretreatment of the sample with H_2O_2 to remove org. matter is omitted. The sp. gr. of the original samples and the mech. fractions of the various types of phosphate rock produced in the U. S. do not vary sufficiently to preclude the use of an av. sp. gr. of 3.0 in calcg. the settling velocities of phosphate particles by means of Stokes formula. An exptl. study of the settling velocities of 5- μ phosphate particles at various temps. showed that the rate of fall is considerably slower than the theoretical rate calcd. by means of the customary form of Stokes formula, this is due apparently to the irregular shape of the phosphate particles. The particle-shape factor in Stokes formula as applied to natural phosphates was found to be 0.154. Results obtained by the pipet method, as modified for the analysis of finely divided phosphates, were in excellent agreement with those obtained when 500-g. samples of the phosphates were sepd. into mech. fractions by sedimentation in water regardless of any theoretical considerations, the progress and efficiency of the sepn. being detd. solely by microscopical examn. of the fractions. The possibility of utilizing the pipet method in the mech. analysis of ground limestone and marls is pointed out. W. H. ROSS

Behavior of tricalcium phosphate in cultivated soils. H. DANNEEL AND K. W. FRÖHLICH. *Landw. Jahrb.* 72, 51-8 (1930).— $\text{Ca}_3(\text{PO}_4)_2$ does not dissolve in water in the mol. form. The ratio $\text{CaO}:\text{P}_2\text{O}_5$ in soln. varied from 1:1.1 at dilns. of 1 in 200 to 1:2.6 at dilns. of 1 in 50,000. The bearing of this on the compn. and soly. of naturally occurring phosphates is discussed. B. C. A.

Technical problems of the superphosphate industry. BRUNO WAESER. *Metallbörse* 21, 75-6, 123-4, 173-4 (1931).—If crude phosphate meal is stirred with sufficient dil. H_2SO_4 to form CaHPO_4 and convert CaCO_3 , Fe_2O_3 and Al_2O_3 into sulfates, a part of the impurities is sepd. in the acid filtrate and the residue contains CaHPO_4 , CaSO_4 and undecompd. CaF_2 and silicic acid. CaHPO_4 is then changed to $\text{CaH}_2(\text{PO}_4)_2$ with more H_2SO_4 . The prepn. of H_2O -free $\text{CaH}_2(\text{PO}_4)_2$ is technically possible by the treat-

(1930)—A humus-contg. product sold under the trade name "Hummit" was found practically valueless from the standpoint of crop production. H. B. SIKES

The solubility of water-insoluble phosphates in silicic acid. W. OEST. *Centr. Zuckerind.* 39, 67 (1931)—A discussion of the effect of Thomas' slag meal (fertilizer) J. F. LESTER

Decomposition of organic fertilizers. II. Relations between the chemical composition and decomposition. SIGBERT OUSSET AND SURIN YOSSE. *J. Ag. Chem. Soc. Japan.* 6, 917-20 (1930)—Soy-bean oil cake, rape-seed oil cake, *Astragalus sinensis*, Sautwickien, rice straw, herring cake and bone fertilizer were analyzed. Seventy-five g. of these org. substances mixed with 1.5 kg. soil were kept in the pots (the area is 1/50000 ha) under the condition that max. moisture content is 50% and temp. 25°. The decompn. velocity was rapid during the first 17 days. The ether ext. is easily decomposed. The decompn. of an aq. ext., especially of reducing sugars and org. N compds., is generally easy. Water-sol. ash of animal origin becomes insol. with greater difficulty than does that of plant origin. The decomps. of cellulose and hemicellulose are, resp., 70-80% and 60-70% in the presence of abundant N. The decompn. of org. N is under 60%. This may be due to absorption by microorganisms. The decompn. of lignin is difficult. Y. KURIARA

Effect of barn manure on soil conditions. S. GOV. *Z. Pflanzenkrankh. Düngeung u. Bodenk.* 9B, 431-7 (1930)—Soils to which barn manure was applied showed, after a year, a slight increase in hydrolytic soil acidity. Manured plots usually showed, after removal of the crop, more residual root-sol. P_2O_5 and K_2O than did the check plots as shown by the Nembauer seedling method. Although the manure added more N than the plants absorbed, there was a definite loss of N as shown by K'ning's method. H. B. SIKES

The influence of drying manure on the nitrogen losses and on crop yields. I. P. MANCHENKOV AND I. F. ROMASHENKOV. *Udobrenia i Urozhai (Fertilizers and Crops)* 2, 294-7 (1930)—Manure was taken to the field and either plowed under immediately or after 1, 2 and 5 days. The manure which was not to be plowed under immediately was placed in the field either in piles or scattered. It was found that the manure became almost air-dry. Analyses showed that the manure in the piles lost more N than the scattered manure. In the scattered manure the losses consisted primarily of NH_3 , which escapes during the first 2 days. Expts. with oats have shown that the manure which was allowed to dry out in the field was slightly inferior. In pot expts. both fresh and dried manure gave the same results. When lime was added to the manures the dry manure was not so effective. J. S. JOFFE

A comparison of the various methods of storing manure. I. P. MANCHENKOV. *Udobrenia i Urozhai (Fertilizers and Crops)* 2, 254-91 (1930).—Three methods were tried: (1) loose throwing together of manure, (2) compacting, (3) the Krantz method (cf. C. A. 21, 3-53). In all cases the manure was cow manure with the addn. of 12% fresh horse manure. The various batches were kept in specially prepd. pits. In the 1st case the manure was piled up loosely 1.25 m. deep and kept that way. For the Krantz method the manure was piled up to 1 m., allowed to heat up to 55° and then compacted. After that another layer 1 m. deep was placed and again allowed to heat up to 55° and compacted. Two series were run: one with straw, the other with peat-moss as bedding. Detns. were made for NH_3 , N on a water ext., then the manure was treated with MgO to det. the combined N. Besides that, a sep. sample was treated with 0.05 N HCl and the ext. distd. over for total NH_3 . In the fresh manure there was found, on the basis of 100 parts of abs. dry wt., 2.7% total N, of which 1.773% was protein N, 0.818% NH_3 -N and the rest other forms. This was for the manure with straw. For the manure with peat the following was found: total N 3.3, protein N 1.877 and NH_3 -N 1.248%. After 4 months the pits were opened and analyses made on the manure batches. The loosely packed manure lost 27.3% dry matter with the straw and 23.0% with the peat, the compacted, 9.9% and 3.4%; the Krantz manure, 21.0% and 10.2%. The N losses were: with method (1) 36.7% with the straw and 34.66% with the peat; method (2) 12.59% and 25.94%; method (3) 26.3% and 19.0%. The distribution of the forms of N in the manure kept by the 3 methods was: in (1) total N 2.35 with the straw and 2.8 with the peat, protein N 1.752 and 2.256, free NH_3 and $(NH_4)_2CO_3$ N 0.345 and 0.152, combined NH_3 -N 0.152 and 0.164, NO_3 -N 0.0015 and 0.009. in (2) total N 2.62 and 2.53, protein N 1.698 and 1.222, free NH_3 and $(NH_4)_2CO_3$ N 0.208 and 0.436, NO_3 -N 0.0015 and 0.0022; in (3) total N 2.520 and 2.97, protein N 1.587 and 1.824, free NH_3 and $(NH_4)_2CO_3$ N 0.415 and 0.829, combined NH_3 -N 0.434 and 0.208, NO_3 -N 0.0018 and 0.0018. Thus in the loosely packed and Krantz manure there is a transformation of mineral N into protein N, but there is a greater loss of N. Pot expts. (5 kg

of soil per pot) were conducted with the variously stored manures on oats. NaNO_3 was used as the check pot. All received similar amts of P and K. It was found that NaNO_3 gave better results than did the manures. The mineral N from the manure is immediately available, whereas the protein N is not. It is concluded that the system of packing down the manure is the best one, since it preserves the mineral forms of N.

J. S. JOFFE

The use of excess molasses. H. CLAASSEN. *Deut. Zuckerind.* 55, 1377-8 (1930).—The author discusses the use of beet molasses as fertilizer and recommends further experimentation.

J. F. LEETE

Floating factories. H. HAUPTVOGEL. *Chem.-Ztg.* 53, 786-7, 806-7 (1929).—A detailed, illustrated description is given of vessels equipped to catch and prep fish and produce fertilizer from the waste. The modern, floating whale oil plants are described briefly.

E. M. SYMMES

Organic fertilizers for oats and flax. Z. V. LOGVINOVA AND A. P. SHCHERBAKOV. *Udobreniia i Urozhai (Fertilizers and Crops)* 2, 476-82 (1930).—N sources from meat scrap, horn meal, horn shavings, burned horn meal, blood meal, slime from intestines, feathers and down, mil meal, tobacco dust and wool combings were compared in pot expts with NaNO_3 and $(\text{NH}_4)_2\text{SO}_4$ on oats and flax. Two sets of pots with 4.5 kg. of soil were set up. One set received 0.5 g. N, the other, 1.0 g. All pots received P in the form of Na_2HPO_4 and K in the form of K_2SO_4 . The results with the oats were: the slime gave almost as good results as the mineral forms of N. The meat scrap, horn meal, horn shavings, burned horn, blood meal, feathers and down and oil meal fell behind the mineral N when 0.5 g. of N was added. With the double quantity of N these materials were just as efficient as the mineral N. The dried blood, tobacco dust and wool combings were far behind the mineral N even when 1.0 g. of N was added. The residual effects of the various org. forms of N on the succeeding crop were far superior to those of the mineral N. With flax the 0.5 g. of org. N was just as good as the mineral N, with the 1.0 g. quantity the org. N was superior to the mineral forms of N. The second crop on the flax pots was oats, and since the flax did not utilize the mineral N, it was effective on the oats giving higher yields than the residual org. N. The quality of the flax was also better with the org. forms of N.

J. S. JOFFE

Fertilizing tomatoes, sweet potatoes and muskmelons in a three-year rotation. J. W. LLOYD. Ill. Agr. Expt. Sta., *Bull.* 364, 18 pp (1931).—Field expts continued for 6 years showed that P is especially important in the production of tomatoes. The most efficient form of P was steamed bone meal. K decreased the yield of melons and was of little or no benefit in tomatoes or sweet corn. Com. N in the form of NaNO_3 or dried blood increased the yield of tomatoes but decreased the yields of melons and corn. $(\text{NH}_4)_2\text{SO}_4$ decreased the yields of all crops. Dried poultry manure increased the yields of tomatoes and corn. Limestone, when used with manure, increased the yields of tomatoes, sweet corn and melons, 7, 14 and 25%, resp., as compared with the use of manure alone. For a 3 year rotation of melons, tomatoes and corn, limestone and manure should be added the first year, steamed bone and dried blood the second year, and no fertilizer is required for the third year.

C. R. FELLEAS

Fertilizing of mildewed vines. HENRI ASTRUC. *Prog. agr. sci.* 94, 499-502 (1930).—A recommends treating vine maladies brought on by a wet season by judicious fertilizing. He favors mixes of org. and chem. fertilizers rather than only the easily sol. chem. fertilizers which are rapidly removed from the root zones. Such mixes should furnish nourishment during the entire vegetating season. A N excess should be avoided, as this overstimulates foliage production and thus encourages mildew growth.

F. W. MARSH

Use of manganese in vegetable greenhouses. I. C. HOFFMAN. Ohio Agr. Expt. Sta., *Bimonthly Bull.* 149, 58-62 (Mar-Apr, 1931).—See C. A. 24, 5102.

C. R. FELLEAS

Results obtained by two winter sprayings and one spring spraying on the vegetation of trees and the quality of fruits. L. CHASSER. *Prog. agr. sci.* 94, 502-5 (1930).—These 3 treatments produced the best results. The first winter spraying was with a mixt. of anthracene oil, formal and schist oil, followed a few days later by lime sulfur. In the spring a Cu arsenical spray was used.

F. W. MARSH

Solubility of Bordeaux. G. L. HOCKENYOS. *Phytopathology* 21, 231-4 (1931).—The colorimetric method of Callan and Henderson (cf. C. A. 24, 312) for detecting Cu, which is sensitive to 1 part in 25,000,000, is applied to the detn. of the percentage of Cu in a series of Bordeaux mixes, contg. various ratios of Cu to hydrated lime. When the Cu:Ca ratio was 1:0.169, 0.04% Cu was present in soln.; with a ratio of 1:0.27 it fell to 0.00009%; with a ratio of 1:0.67 the sol. Cu was found to be 0.00013%, and it remained

at this level when the ratio was made 1 for 1 2. The detn. is made by adding to 17.5 cc. of the soln. to be tested 0.5 cc. NH_4OH and 2 cc. 1% Na diethylthiocarbonate and comparing in a Duboseq colorimeter with 2 cc. 0.0002 N CuSO_4 soln. diluted to 17.5 cc. and similarly treated.

JOSEPH S. CALDWELL

Factors influencing the character of Bordeaux mixture. W. NEWTON, F. B. JOHNSON AND C. YARWOOD. *Proc. Can. Phytopath. Soc.* 1929, 21-6(1930), *Rev. Applied Mycol.* 10, 118.—The ideal spray is probably obtained when the solids are in a state of max. dispersion. Best results and slowest settling is obtained when the reacting solns. are as cold as possible, when the CuSO_4 soln. is poured slowly into the lime suspension rather than conversely and when the CuSO_4 is as dil. and the lime suspension as concd. as practicable. Addn. of Na silicate, fresh skim milk, various alcs., wheat flour, whale oil soap, various washing soaps, agar, alkali, or sodium resin soap failed to improve the Bordeaux mixt., while 0.15 to 0.5% Ca caseinate, 0.5% gelatin, Crystal White brand soap, or potassium resin soap did improve the Bordeaux somewhat.

ODEN E. SHEPARD

Analysis of commercial lime-sulfur materials. TATSUNOSUKE ONOTE, *J. Imp. Agr. Expt. Sta. (Japan)* 1, 89-90(1929)—High sp. gr. of lime-sulfur soln. does not always indicate high sulfide-S content. The quality should be detd. by detn. of active-S content and not by sp. gr.

K. KITSUTA

Determination of barium fluosulfate spray residue. R. H. CARTER, *Ind. Eng. Chem., Anal. Ed.* 3, 146-7(1931)—The apples are washed by immersion for 30 sec. in a boiling 3% NaOH soln. and rinsed with slightly acidulated H_2O . The rinsings and washing soln. are cooled, made strongly acid with HCl and rapidly filtered. The filtrate is then made strongly alk. with NaOH, and the Ba is pptd. as BaSO_4 with an excess of H_2SO_4 . The BaSO_4 is then calcd. to Ba fluosulfate. This method is fairly rapid and gives reproducible results which are sufficiently accurate for comparative work.

J. R. ADAMS

Tests of treatments against the apple worm. PEYKHA, *Prog. agr.* vol. 95, 81-3 (1931)—The treatment of apple trees with Cu-arsenical sprays 3 times during the progress of flower formation proved effective.

F. W. MARSH

The fungicidal action of ultra-violet radiation. W. A. R. DILLON WESTON AND E. T. HALLAN, *Phytopathology* 20, 939-65(1930).—The work in part parallels that of Fulton (cf. C. A. 24, 1926). Cultures on potato agar were irradiated for varying periods with a quartz Hg-vapor (Alpine sun) lamp through covers of Vita or Sanalux glass. The effect was to inhibit the growth of the mycelium, and when radiation ceased, growth was renewed at a rate proportional to the time and distance of the previous radiation. Under weaker dosagr. growth of the mycelium at the surface ceased but persisted deep in the medium. Concordant results were obtained with 12-14 species.

J. S. C.

Physical and chemical action of ultra-violet rays on sublimed sulfur. FONES DIACON, *Prog. Agr.* vol. 95, 155-8(1931)—Expts. showed that ultra-violet rays did not effect the retrogradation of sublimed S. The reason for the action of S on oidium is discussed.

F. W. MARSH

The relation of pentathionic acid and its constituents to the toxicity of sulfur fungicides. O. NEAL LIVING, *Phytopathology* 21, 131-2(1931)—At summer temps. over 10% of the S on a dusted surface may pass off within 2 weeks. S vapor is not toxic to fungus spores. H_2S is present only in traces and is non toxic in those concns. Pentathionic acid is present in concns. toxic to fungus spores, and its formation is enhanced by mild oxidizing agents and possibly by H_2S . It is stable in both acid and weak alk. solns., but toxic only in the former. It is not sufficiently volatile in dil. solns. to be toxic at a distance. The toxic action is governed by the condition of the fungus rather than by that of the pentathionate ion.

JOSEPH S. CALDWELL

Sulfuring apricots. GEO. QUINN, C. G. SAVAGE, A. V. LYON AND W. R. JEWELL, *J. Dept. Agr. S. Australia* 34, 513-4(1931).—The authors, serving as a comm., recommend in detail a procedure for sulfuring fruits.

M. S. ANDERSON

Fumigating tomato houses with hydrocyanic acid gas to destroy white fly (*Trialeurodes vaporarum*). GEO. QUINN, *J. Dept. Agr. S. Australia* 34, 519-20(1931).—A detailed procedure is given for the use of HCN as a fumigant.

M. S. ANDERSON

Mercury in the disinfection of seeds. VITTORIO CASABURI, *Industria chimica* 5, 1474-82(1930), cf. C. A. 25, 1942.—The various Hg compds. previously studied are compared, and the conclusion is that Dry Ton is the most efficient, both as an insecticide and in the promotion of germination.

A. W. CONTIERI

Disinfection experiments on cotton seeds with the dusts Tillant R and ceresan. F. FORSTENREICHNER, *Nachr. über Schädlingbekämpfung* 5, 136-47(1930), *Rev. Applied*

Mycol 10, 103—Excellent control of a species of *Rhizoctonia* and the facultative parasites accompanying it was obtained by use of Tillantin R and cerasan at the rate of 1 kg per 100 kg seed. Cerasan is most effective. The cotton seedlings are infected commonly by this species in the Adana region of Turkey. ODEE E. SHEPPARD

Contribution to the problem of cereal rust control by chemical methods. G. GASSNER AND W. STRAIB. *Phytopath* 2, 361-76(1930). *Rev Applied Mycol*, 10, 88-9—S with and without an equal admixt. of kaolin, 1% Bordeaux, 1% Cu chloride, Ca cyanamide mixed with kaolin, basic slag and calcined lime dusts and kainite were used in the control of brown and yellow rust (*Puccinia triticina* and *P. glumarum*) on Heine's Kolben and Dickkopf wheat and of *P. dispersa* (*P. secalina*) on Petkus winter rye. Ca cyanamide mixed with large quantities (up to 90%) of kaolin gave very excellent protection, and G. and S. consider its use on a large scale. S and Bordeaux gave very good control, but the others gave little control, while Cu chloride caused much leaf burning, and NaNO_2 actually increased the rusts. ODEE E. SHEPPARD

The treatment of the rot of wheat (1929-1930). G. ARNAUD AND MISS GAUDINEAU. *Compt rend acad agr France* 16, 1029-35(1930), cf *C A* 24, 4577.—Among the anti-cryptogamic chemicals a 5% CuSO_4 soln. is very satisfactory. The use of the CuSO_4 soln. is followed by a dusting of the grains with CaO which in turn is followed by a thorough washing with H_2O . The Bordeaux mixt.-casein prepn. has given some excellent results, and in many cases treatment with Iormol has been very efficacious. J. R. ADAMS

Effect of some factors influencing disease conditions of potatoes. H. REILING. *Z Pflanzenkrankh Düngung u Bodenb* 9B, 393-7(1930)—Potato scab occurs principally on alk. soils. Acid soils usually produce potatoes free from scab, but tubers and the plant itself may become infected with *Rhizoctonia solani*, particularly during drought. Lack of potash in the soil may produce brown rings in the tuber and on the surface. H. B. SIEMS

The leaf diagnosis of the potato. L. HENRI LAGATU AND LOUIS MAUMS. *Ann. sci. agron* 47, 595-653(1930)—The chem. state of a leaf taken from a conveniently located place at a particular moment is called "the leaf diagnosis at a given instant." "The annual leaf diagnosis" is the series of chem. states of the leaf revealed by the analyses at different periods during the entire growing cycle. The results of this method are compared with indications furnished by the yields and by observations on developments. The percentages of potash, N, acid phosphate and lime detn. by the leaf diagnosis method on samples harvested on 5 dates in May and June from plots differently fertilized showed clear variations. Plots similarly fertilized gave concordant results. A definite relation is established between the leaf diagnosis results and the fertilizers used. In general, N and acid phosphate in the leaf decreased with plant age, while lime increased. Comparisons of diagrams for N and potash show a rise of the N curve, or a lowering of that for potash or the two displacements together. Conclusion N was needed for potatoes, but not acid phosphate or potash. In another expt. an excess of N in certain complete fertilizers was observed. Results of the study of the effect of the conditions of the cultivation media on the plant and on the leaf composition are considered. F. W. MARSH

Effect of seed-potato treatment on yield and rhizoctonosis in Florida from 1924 to 1929. L. O. GRATZ. *Fla. Agr. Expt. Sta., Bull* 220, 5-30(1930)—Numerous field expts. conducted over a period of 6 years show no yield increases and pos. indications of definite injury from the use of org. Hg compds. for seed potato treatment. The use of HgCl_2 , hot CaH_2O or org. mercurials is not justified under Fla. conditions. Ninety two references are appended. C. R. FELLERS

Increased yields from spraying and dusting late potatoes. J. E. HOWITT. *Proc. Can. Phytopath. Soc.* 1929, 19-20(1930). *Rev Applied Mycol* 10, 125-6—In 1928 potato plots treated with liquid Bordeaux (4-8-40) yielded 291 bushels per acre, plots dusted with copper lime yielded 235 5 bushels and control plots 194 5 bushels, while in 1929 the corresponding figures were 235, 193 and 156. In 1928 the beneficial effect was probably due to the preventive action of the fungicides on late blight and rot, and in 1929 the beneficial effect was correlated with tip burn. ODEE E. SHEPPARD

The control of oidium leaf disease. R. K. S. MURRAY. *Trop Agr. (Ceylon)* 75, 294-302(1930)—S. dusting is a quick and effective means of controlling oidium leaf disease. JOHN O. HARDESTY

Observations concerning the development and struggle against mildew in 1930 in the Isère. L. ROUCIER. *Prog. agr. vit* 94, 566-9(1930).—Cu was found to be the best remedy for use against this parasite. F. W. MARSH

Copper and mildew. JACQUES LUGAN. *Prog. agr. vit* 94, 452-3(1930)—L. dis-

agrees with Villedieu (C. A. 25, 372) in certain conclusions concerning the use of ammoniacal fertilizers for grapevines and of Cu as a mildew combatant. F. W. MARSH

The effect of boron on powdery mildew and spot blotch of barley. FRANK M. EATON. *Phytopathology* 20, 967-72 (1930).—Barley was grown in pans of quartz sand, so arranged as to be flushed daily with nutrient solns. In crops grown during the summer, the plants receiving no B were heavily infested with mildew but were free of spot blotch, those receiving 5-25 p p m of B were free of mildew but showed heavy infestation with spot blotch. In winter plantings, mildew was present in all the beds, most abundantly in those receiving no B, and decreased in amt with increase in B applied. Spot blotch did not appear in the B free bed but was abundant in all those receiving B. On wheat, powdery mildew was present irrespective of the amt. of B supplied. The B content of the summer-grown plants ranged from 57 p p m dry weight on the B-free cultures to 1349 p p m in those receiving 25 p p m of B in the culture soln.

JOSEPH S. CALDWELL

Lessening of growth of rice plants. MIRANDA LANZA. *Boll. lab. sper. fitopatol.* (Torino) 8, No. 1, 6-9 (1931).—The observed disease is due to fungous parasites which winter on the seeds. They may be destroyed by immersion for 24 hrs. in 0.2% Kalimat (a mixt. of formaldehyde and phenol).

G. A. BRAVO

Sugar beet root-rot control. H. C. YOUNG. Ohio Agr. Expt. Sta., *Bimonthly Bull.* 148, 15-8 (Jan. Feb., 1931).—Field expts. showed that the application of 500 lb. of NaCl per acre often materially increased the stand of beets, and no injury to the crop was recorded on any plot.

C. R. FELLERS

Use of oxides of unsaturated hydrocarbons for the eradication of barberries and other pests. R. B. HARVEY. *Phytopathology* 21, 126 (1931).—Ethylene oxide and propylene oxide have been used for killing gooseberries or barberries. A depth charge is injected beneath the plant by means of a prod rod or "gopher stick" provided with a measuring chamber and connected directly with the gas tank, or a soln. of ethylene oxide in water can be injected from a knapsack sprayer. The stems blacken and the leaves fall off, the plants dying in 1-2 weeks.

JOSEPH S. CALDWELL

Pyrethrum (ZOCHAREWICZ) 17. Fodder and quality of milk in meadow fertilization with CaCN_2 (HERMANN) 12. A brief history of the world's phosphate rock production (GRAY) 18. Tung-oil tree (NEWELL, *et al.*) 26. Mineralogical study of the soil of the Vercelli district (REPOSSI) 8. Lignin, humic acid and humin (FUCSIS) 10. Influence of fertilizer treatment on the vitamin A content of spinach (HONEYWELL, DUTCHER) 11E. I content of Lettish waters in relation to the improvement of crops (KUPCIS) 14. Cyanamides [fertilizer] (Brit. pat. 338,023) 18. Discharge device [apparatus for atomizing insecticides with CO_2] (U. S. pat. 1,794,185) 1.

Handbuch der Pflanzenernährung und Düngerlehre. Edited by FRANZ HONCAMP. Band II. Düngemittel und Düngung. Contributions by Ernst Hugo Birei, Hand Brenck, R. Demoll, *et al.* Berlin: J. Springer. 919 pp. M. 86, linen, M. 89 80.

Fertilizer. OTTO RUPKE (to I. G. Farbenind. A.-G.). U. S. 1,793,133, Feb. 17. NH_4CNS is heated (sustably at a temp. of about 300°) until evolution of gases has ceased.

Fertilizers. A. HOLZ and T. VAN D. BERDELL. Brit. 337,402, July 30, 1929. Tri Ca phosphate is decomposed with mixed H_2SO_4 and HCl, and the insol. CaSO_4 is sepd. from the sol. H_2PO_4 and CaCl_2 . The CaSO_4 is treated with a K compd. such as KOH in the presence of CO_2 (preferably under pressure and heat) or K_2CO_3 to form K_2SO_4 and CaCO_3 . The CaCO_3 is sepd. and is added to the soln. contg. H_2PO_4 and CaCl_2 in the presence of $\text{Ca}(\text{OH})_2$ to ppt. di Ca phosphate, which is then sepd. and mixed with the K_2SO_4 soln., and the mixt. is spray-desiccated to produce a granular product. Various auxiliary procedures are described.

Fertilizer. A. HOLZ and T. VAN D. BERDELL. Brit. 337,415, July 30, 1929. A mixed fertilizer free from chlorides is prepd. by treating KCl with H_2SO_4 to form KHSO_4 and HCl gas, expelling the HCl (preferably at a temp. well below 200°) and then treating the KHSO_4 with NH_3 to form K and NH_4 sulfates. The HCl is passed into an aq. suspension of rock phosphate to form CaCl_2 and H_2PO_4 , insol. solids are sepd. and sufficient $\text{Ca}(\text{OH})_2$ is added to ppt. di-Ca phosphate, which is then sepd. and dried (preferably at a temp. not much above 100°) and mixed with the K and NH_4 sulfates. Various details and slight modifications of procedure are described.

Fertilizers. 1 G FARSEVIND A-G (Ernst Jänecke, inventor). Ger. 514,509, Nov 20, 1929 Nitrogenous fertilizers are prepd. by leading gases contg N_2O_4 into a soln. of NH_3 salt in concd. H_2SO_4 and HNO_3 , and working the product up into a nitrate- and sulfate-contg product. In the example, N_2O_4 is led into a soln. of $(NH_4)_2SO_4$ in 50% HNO_3 to give the fertilizer $(NH_4)_2SO_4 \cdot NO_3$.

Fertilizers. N V MAATSCHAPPIJ TOT EXPLOITATIE VAN VEREDFLINGSPROCÉDÉS. (Paul Erasmus, inventor) Ger. 514,510, Nov. 7, 1928 Carbohydrates of vegetable origin such as peat, lignite, sugar, cellulose, wood, etc., are heated to about 300° with NH_3 under pressure to form a nitrogenous fertilizer. Examples are given.

Fertilizers. KUNSDÜNGER-PATENT-VERWESTUNGS A G Swiss 141,328, Oct. 16, 1928. See Brit. 310,276 (C A 24, 679)

Fertilizers. SOCIÉTÉ D'ÉTUDES CHIMIQUES POUR L'INDUSTRIE Swiss 140,729, Nov 20, 1929 A nitrogenous fertilizer contg $CaCN_2$ is produced by the action of NH_3 on $CaCO_3$ in a current of mixed H and N at $500-1000^\circ$. Cf C A 25, 767.

Fertilizer and sodium bicarbonate. SOC. CHIMIQUE DE LA GRANDE PAROISSE AZOTE ET PRODUITS CHIMIQUES Brit 338,007, Jan 11, 1929 The process for producing a composite fertilizer and $NaHCO_3$ by double decompn. of $NaNO_3$ and KCl and treatment of the mixt produced with NH_3 , CO_2 and water, as described in Brit. 331,451 (C A 25, 165), is modified by effecting the double decompn. reaction of $NaNO_3$ with the KCl in the mother liquors of the process, as by introducing these materials into the soln. at that point of the manuf. at which the pptn. and sepn. of $NaHCO_3$ have been effected.

Rendering insoluble phosphates available for fertilizer. HANS LEIRACKER U. S. 1,795,173, March 3 Material such as mineral phosphate is heated to about $1000-1400^\circ$ with such quantities of C and silica-contg addns. that part of the total P content of the initial material is liberated in accord with the equation: $P_2O_5 + 5C = P_4 + 5CO$ and that in the calcined residue contg the remaining P there is a proportion of P_2O_5 about 0.9-1.1, SiO_2 about 0.9-1.1 and CaO about 4.8-5.2 mols.

Treating phosphate rock. KUNSDÜNGER-PATENT-VERWESTUNGS A - G Brit. 338,079, March 7, 1929. In leaching phosphate rock with HNO_3 and K_2SO_4 a part of the soln. of H_2PO_4 and KNO_3 from a previous operation (partly freed from KNO_3 by cooling and crystn.) is used as a diluent, thus preventing the pptn. of $CaSO_4$ contg a high percentage of K. Various details of treatment for prep. fertilizer are described, which may include addn. of NH_3 .

Superphosphates. FRANCESCO C. PALAZZO and FORTUNATO PALAZZO Swiss 141,501, Nov 3, 1928 Superphosphates with 42-51% of P_2O_5 capable of assimilation are obtained by repeatedly extg a superphosphate contg 14-17% P_2O_5 capable of assimilation with half its weight of water, pptg the H_2PO_4 by milk of lime and adding the $CaHPO_4$ to other exts contg free H_2PO_4 . Cf C A 24, 680

Superphosphates. FRANCESCO C. PALAZZO Swiss 142,459, Sept. 27, 1928 Phosphates with a content of 42-45% assimilable P_2O_5 are prepd. by kneading concd. H_2PO_4 (e.g., obtained by working up phosphorites) with pure pptd. $CaHPO_4$ at $40-50^\circ$ for about 2 hrs. The mixt. is vacuum dried and powd.

Superphosphate chambers. D. WALLISCH Brit. 338,075, Jan 31, 1930 Various different constructions are described designed so that the reaction mass and its support are moved in relation to that part of the chamber which contains the charging device. One form comprises a cylindrical chamber mounted for rotation around a vertical axis and which, after initial charging, is adapted for continuous operation. Knives or compressed air may be used to break up the mass.

Apparatus for leaching out raw materials. KUNSDÜNGER-PATENT-VERWESTUNGS A - G Swiss 141,300, Oct. 16, 1928 Raw material is placed in a succession of tanks of which each is lower than the preceding one, and the acid solvent from the bottom of one tank is led to the top of the next higher tank, while the overflow of the higher tank drips into the top of the lower tank. The method is suitable for the prepn. of artificial fertilizer contg sulfates, carbonates, etc.

Jacketed sectional drying apparatus for preparing fish meal. HAROLD S ROWTON and VOLKMAR W HÄNIG U S 1,793,638, Feb 24 Structural features.

Forcing plants. FARBEALZ C M N H Ger. 517,115, Nov 25, 1928 Unstable cyanides giving rise to HCN and another forcing agent, e.g., NH_3 or C_2H_2 , are used $Ca(CNNH_2)_2$ and crude $Ca(CN)_2$ contg free CaC_2 are suitable, the former being obtained by prep. $Ca(CN)_2$ in liquid NH_3 .

Insecticide. HANS WESCH, KARL BRODERSEN and WERNER EXT (to Winthrop Chemical Co) U. S. 1,794,046, Feb 24 An aromatic monothiocyanate such as benzene thiocyanate is used as an active constituent of preps. such as powders or solns.

Applying insecticides and fungicides to growing plants, etc. JOSEPH D. NAULA.

U. S. 1,702,860, Feb. 17. A porous carrier such as sandust or pumice impregnated with a mineral oil is subjected to simultaneous distill and incomplete combustion, and the vapors thus produced are projected on trees or plants or other articles for their insecticidal and fungicidal effect. App. is described.

Seed-goods protection. VERGIN FOR CHEMISCHE UND METALLURGISCHE PRODUKTION. Ger. 514,681, July 2, 1921. A prepn for protecting seed goods from vegetal and animal blight contains sol. complex compds. or mixts. of nuclear metallized aroma compds., with org. substitution products of NH_3 . In the example, HgO is dissolved in H_2SO_4 and heated with cresol on a water bath. NaOH is added, and the ppt. dissolves in piperidine. Or pyridine and KCNS is heated with a Cu salt and the product mixed with ethylene diamine.

Seed-disinfectant composition. FORST J. LUNK (to E. I. du Pont de Nemours & Co.). U. S. 1,701,680, Feb. 21. HgCl_2 is used in admixt. with a finely divided mineral such as bentonite to form a prepn. capable of producing relatively stable and substantially homogeneous suspensions with water.

Seed-disinfecting composition. FORST J. LUNK (to E. I. du Pont de Nemours & Co.). U. S. 1,701,622 Feb. 21. Mercurized o-chlorophenol and mercurized o-nitrophenol are used with bentonite. Other similar mixts. also are described.

Facilitating germination of seeds. WARRIN TILDEN & CO. Brit. 337,601, Nov. 15, 1928. See U. S. 1,760,188 (C. A. 24, 3077).

16 THE FERMENTATION INDUSTRIES

C. N. FRAY

Annual analytical report. B. LAMPE. *Z. Spiritusind.* 53, 281-2(1930).—Report on the activity of the technical scientific laboratories of the German alcohol manufacturers' association and the association of potato dryers.

Are asbestos-slate-lined fermentation tanks suitable for fermentation? W. KR. *Z. Spiritusind.* 53, 268(1930). The answer is in the negative.

Absolute alcohol. F. WAGNER. *Z. Spiritusind.* 53, 260-1(1930).—A brief review.

Improvements in the manufacture of absolute alcohol. H. GUINOT. *Chim. et Industrie* 25, 26-32(1931).—A description is given of the evolution of so-called azeotropic methods, which now permit of the production of pure absolute alcohol directly from fermented wort by the use of the standard equipment at present used for the production of rectified 96% 100% and at but slightly higher cost. A. PAPINEAU COURTEAU.

New process of drying potatoes and the manufacture of alcohol therefrom. L. OUMER. *Z. Spiritusind.* 53, 251-3(1930), cf. C. A. 25, 1327.—An economic discussion.

Evaluation of potato flakes according to their yield of alcohol. B. LAMPE. *Spiritusind.* 53, 274-5(1930).—There is no relationship between moisture content of potato flakes and yield of ale. It is necessary to score the sample for outside appearance, color, odor, uniformity, development of bacteria, moisture content, yield of ale, starch value. Correct evaluation can be computed with these results on hand.

Composition of compressed and dried potato crumbs. B. LAMPE. *Z. Spiritusind.* 53, 276-7(1930).—The composition of potato juice was given in an earlier paper (C. A. 1327). The analytical results show that the pressed and dried potato crumbs contain about 40-50% of the ash usually present in potato flakes. The same relation is found for protein.

Effects of mashing methods and souring on potato flakes. B. LAMPE and W. KR. *Z. Spiritusind.* 53, 277-8(1930).—Two different processes were tried. The necessary proportion of malt was added in 2 portions in 1 expt. and mashed in one operation in another expt. The yield of ale was smaller and the steam consumption was greater in case the malt was added in 2 portions, provided the mash was vigorously stirred and not allowed to form lumps. Artificial souring diminished the yield of ale.

Addition of ammonium salts to vinegar. C. ARNSWORTH MITCHELL. *Analyst* 178-9(1931).—Analyses show that the practice has been revived of adding NH_4 salts to artificial vinegar to make it contain as much N as when made from malted or malted grain.

Studies on the biochemistry of zinc. Zinc and yeast fermentation. AS. ZLATOFF, M. ANDRETSCHIEVA AND D. KAITSCHIEVA. *Biochem. Z.* 231, 123-31(1931). Studies made on *Saccharomyces ellipsoidens* show that $\text{Zn}(\text{NO}_3)_2$ and ZnSO_4 have

and Seusslitz A. HEIDUSCHKEA AND C. PYRIKI *Z. Untersuch. Lebensm.* 59, 613-5 (1930), cf. *C. A.* 24, 1699—The range in d of the must was 60.7-88.4° Oechsle, titratable acids were 4.4-12.8 g per l, and mineral constituents 2.64-5.04 g per l. The d averaged 4.2° higher than the av for 1928.

Chianti and its wines. LUIGI RICASOLI *Atti III congresso naz. chim. pura applicata* 1930, 285-91—Descriptive. C. R. FELLERS
C. M. SYMMES

A new category of abnormal wines. CHARLES BERTIN *Prog. agr. vit.* 94, 564-6 (1930)—It is noted that certain normal wines do not correspond to the enological laws concerning the relative amts of alc. and acid that should be present. P. W. M.

The influence of adjuncts on the composition of beer. E. JALOWETZ *Brau- und Malzindustrie* 24, 1-2 (1931)—A discussion. S. LAUFER

The formation of layers in beer in storage vats. E. JALOWETZ *Brau- und Malzindustrie* 24, 21-5 (1931)—W. Kusenak (*Tagesitz. Brauer* 27, No. 175, 28, No. 110) found that layers are formed in beer kept in storage after 4-5 weeks; the layers differ in CO₂, ext. and albumin content, viscosity, acidity and pH value. J.'s investigations agreed fairly well with K.'s findings. Methods are given for maintaining uniform compn in stored beer. S. LAUFER

Preventing the "turning" of beer. J. RAUX *Brasserie & malterie* 20, 342-5, 359-62 (1931)—An address. A. L'APINEAU-COUTURE

The extractives of whiskey. WM. PARTRIDGE *Analyst* 56, 177-8 (1931)—The U. S. Pharmacopoeia states that whisky should not leave more than 0.1 g. of solids when 20 cc. is evaporated. Various English books state that whisky should not contain more than 1/4 as much solids, which is probably attributable to a statement made in 1879 in Allen's *Commercial Organic Analysis* which does not hold true with respect to the whisky of today. W. T. H.

Simple universal thermostat for laboratory fermentations (FINK) 1. [Fermentation] of the aqueous extracts from olives (CIUSA, *et al.*) 27. Detection of iso-PROH in brandy and spirits (REIS) 7. Model experiments for the theory of alcoholic fermentation (OILIE, COURTICOS) 10. Peroxidase preparations (Brit. pat. 337,405) 11A.

TATE, FRANCIS G. H. *Alcoholometry: An account of the British Method of Alcoholic Strength Determination, with an historical introduction written by the author in collaboration with George H. Gabb* London. H. M. Stationery Office, 93 pp. 5s., net. Reviewed in *Nature* 127, 398 (1931).

Device for atomizing fatty substances on fermenting liquid to keep down foam. ARNE BOYE. U. S. 1,792,903, Feb. 17. Structural features of a device designed to operate automatically according to the amt. of foam formed.

Testing saturation of barley, etc. W. & T. AVERY, LTD., W. A. BENTON and P. E. GORF. Brit. 338,282, Sept. 26, 1929. For detg. the degree of satn. of material, such as that of barley during steeping prior to malting, a sample is centrifuged for a definite time, weighed, re-immersed, re-centrifuged and re-weighed, and these operations are repeated until 2 consecutive weighings are the same (indicating that satn. has been attained). App. is described.

Extracting the bouquet from wines and spirits. OTTO FODOR. Ger. 517,061, May 7, 1927. The constituents imparting the bouquet to wines and spirits are obtained practically free from EtOH by treating the beverages, or the mashes from which they are prepd., with adsorbents such as active charcoal or silica gel after or while reducing the alc. content, e. g., by distn. The adsorption may be assisted by adding a little fatty or paraffin oil to the beverage. The bouquet substances may be recovered from the adsorbent by extn. and isolated by evapn. of the ext., and the beverage may be reconstituted by dilg. the bouquet substances with EtOH and H₂O. Examples are given.

Yeast. C. LANGEMAYER. Brit. 337,947, Oct. 24, 1928. A yeast obtained by an otherwise usual process is treated with a small quantity of K salts and albuminous substances, such as KCl or K₂SO₄, and albuminous substances derived from peanut meal, malt germ, lupins or yeast-press liquor, allowed to stand for several hrs. and then pressed. Use may be made of a must of peanut meal, yeast and malt germ, digested with H₂SO₄, H₃PO₄, or lactic acid, with addn. of a small quantity of dextrose.

Hydrolyzed protein products. CHARLES N. FREY (to Standard Brands, Inc.) U. S. 1,794,370, March 3. See Can. 308,910 (*C. A.* 25, 1922).

17-PHARMACEUTICAL CHEMISTRY

W O EMERY

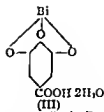
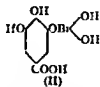
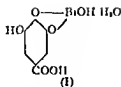
"Enfleurage" of Jassin. TRARALD *Parfumerie moderne* Jan, 1930, *Masloboino Zhivooe Delo* 1930, No. 2, 84-5—In the enfleurage or extn. by fat a high percentage of indole is obtained whereby the glucosides are split. Methyl anthranilate is also obtained; it is absent in the oil from flowers. These constituents improve the odor. One thousand kg. of flowers yields on extn. by fats 5-6 kg. of an ointment which contains 25-28% of a pure flower oil or about 1.2-1.6 kg. while the usual extn. yields 1.3-1.4 kg. of an absolute essence, the latter contg. about 30-40% pure flower oil. The enfleurage method produces a yield 2-2.5 higher than the ordinary extn. methods. Flowers left after enfleurage upon being extd. produce more oil but of a different character. A. A. BOEHLINGE

Mechanical processes for lemon essence extraction. FRANCESCO LA FACE. *Boll. soc. sp. e. essenze* 5, 129-37, 149-50, 178-81, 212-8, 234-43 (1930)—The commonly employed processes for the extn. of the essence from lemon peel are: sponge process (I), mech. rasure processes with the Cannavò (II), Speciale (III) and Calabrese (IV) machines, mech. pressing processes with Perstener (V) and Bennet (VI) machines and a few others. Analyses were made on the samples of each of these essences with the Wilson and Young method. With the IV method there is a max. loss of essence that remains in the fruits (30.5-46%), with the I an av. of 27.5%, and with the II and III a min. loss of 10-11%. The characters of the various types of essence, the working costs and the total yields are very different. I and VI give a yellow essence, V colorless, the others green, the III essence contains 4-4.9% of fixed residue and the V only 0.3-0.65%. The mech. processes give an essence of grass odor but less odorous (the extn. loss may be 25% in regard to the citral content of the sponge essence). Among the causes that may lower the citral content, the most important are the fermentations, therefore it is better to sterilize the extn. water by boiling or by adding Na benzoate. G. A. BRAVO

Colorimetric determination of adrenaline. ITALO SANSONI. *Officina* 3, 323-34 (1930)—Adrenaline can be detd. colorimetrically in pharmaceutical solns. by using standard pyrocatechol solns. with one of the following procedures. (A) Prep. standard solns. with increasing quantities of pyrocatechol, add 5 cc. of 40% Na acetate soln. and 4 drops of 1% FeCl_3 to 5 cc. of each standard and of sample. Compare the resulting colors in test tubes or in the colorimeter in the ordinary way. Since the standard 0.08% pyrocatechol soln. and the 0.1% adrenaline soln. give identical colors, the adrenaline content of the unknown soln. can be easily calcd. (B) To 2.5 cc. of each standard (prepd. with pyrocatechol) and of sample add 2 cc. of Folin reagent (10 g. Na tungstate, 8 cc. of 85% H_3PO_4 and 75 cc. H_2O are heated for 2 hrs. with a reflux condenser, and, after cooling, dild. to 100 cc. and filtered if necessary), 5 cc. of water and 4 cc. of 20% Na_2CO_3 soln., keep all the liquids for $\frac{1}{2}$ hr. at 30° and then compare. The standard 0.14% pyrocatechol soln. and the 0.1% adrenaline soln. give identical colors. (C) To 5 cc. of each standard (prepd. with pyrocatechol) and of sample add 2 cc. sulfanilic acid (4 g. acid and 90 cc. of 37% HCl dild. to 1000 cc.), 0.5 cc. of 1% NaNO_2 soln. and 1 cc. of 28% NH_4OH and then compare. This reaction is the easiest and the most sensitive. The standards have good keeping qualities and give identical colors with adrenaline solns. of the same concn. The ordinary stabilizers added to the pharmaceutical solns. of adrenaline (chlorotone, physiol. NaCl soln., etc.) do not interfere. G. A. BRAVO

The methods of titrating the alkaloids in the Solanaceae. ENRICO DE CECCO. *Officina* 3, 335-42 (1930)—The Italian Pharm., 5th Ed., suggests titrating the alkaloids in the drugs by the method of titrating the alkaloids in the drug exts. With this procedure, great errors may result from the decomposition of the alkaloids in aq. soln. and from the desiccation of the residues. Comparative expts. were made with German, English and U. S. P. on stramonium and belladonna. The U. S. P. method is the best, but the results are more accurate and very concordant if 0.01 N solns. are used instead of 0.1 N solns. in the titrations. G. A. BRAVO

The constitution of dermatol. ANTONIO GRIPPA. *Officina* 3, 343-51 (1930)—G. did not succeed in prep. the Me, Ac and Bz derivs. of dermatol., therefore the hydroxyl groups do not exist free in the mol., and the constitutions I (cf. Prunier, *J. de pharm.* 2, 498 (1901)) and II (cf. Biétrix, *Bull. soc. chim.* 1, 692 (1893)) are not exact. Moreover, dermatol., when heated at 110° , loses 2 mols. of H_2O . The constitution III (cf. Caseneuve, *Bull. soc. chim.* 1, 852 (1893)) is proposed for dermatol., as prep. by the method of the Italian Pharm., 5th Ed.



G. A. BRAVO

Extraction of orange oil F. J. DE VILLIERS. *J. S. African Chem. Inst.* 13, 48-52 (1930).—Steam distn. of orange peel causes some decompn. of the essential oil unless carried out in high vacuum. Ale. proved the best solvent for extn., the peel being rasped under ale. and prolonged contact being avoided. Addn. of water to the ext. causes the sepn. of a fine emulsion difficult to sep. directly by centrifuging. The particles of the emulsion are negatively charged and sepn. is effected by the addn. of electrolytes. The absorption of ions by the emulsion was in the order $\text{Fe} > \text{Ca} > \text{Na}$, but the most effective coagulant was the H^+ ion. B. C. A.

Scheme for the microchemical identification of alkaloids. J. F. H. AMELINK. *Pharm. Weekblad* 68, 159-55 (1931).—The scheme consists in treating the sample with a series of alkaloidal reagents and noting the character and microscopic appearance of the reaction products. The reagents employed are PtCl_4 , $\text{PtCl}_4 + \text{NaI}$, AuCl_3 , $\text{AuCl}_3 + \text{NaBr}$, HgCl_2 , $\text{K}_4\text{Fe}(\text{CN})_6$, $\text{K}_3\text{I}(\text{CN})_6$, Dragendorff's reagent (K_2BiI_4), KOH and picric acid. The microtechnique for extn., filtration and sublimation of the alkaloid is described in some detail. The reactions are tabulated and a large no. of drawings are given showing the typical cryst. forms of the reaction products. In some cases the alkaloid can be identified by a single reaction, but usually 2 or more reactions are required for identification. In special cases it is desirable to make confirmatory tests by specific reactions. The table gives 78 substances, mostly alkaloids, but includes also other products commonly classed with the alkaloids, e. g., caffeine, cantharidin, santonin, simple bases, e. g., betaine, choline, synthetic medicinals, e. g., PhNHAc , antipyrine, urotropine. *Ibid.* 211-6. The reactions are discussed in detail as applied to *percarine*, α -butoxy γ -(diethylamino)thylcarboxamidoquinoline, and *diocaine*, *di-N,N'-(p-allyloxyphenyl)acetamidine*, 2 local anesthetics included in the above table. A. W. DOX.

Production of new disinfection preparations of thymol and carvacrol. I. Production of thymol and carvacrol. C. PHILIPP AND PH. KUHN. *Pharm. Presse, Wiss.-prakt. Hefte* 1931, 19.—A brief survey of the principles upon which the prepn. of cymene, thymol, carvacrol and mono- and dichlorocarvacrol is based. II. Disinfectant action of thymol and carvacrol preparations. PH. KUHN. *Ibid.*—The results obtained in a series of expts. on the disinfectant action of thymol, carvacrol and their Cl substitution products are reported. The most valuable of the products tested proved to be chlorocarvacrol (carvasept) which can be usefully employed where phenol and thymol are indicated. Its advantages lie in a higher disinfectant power, lower toxicity than PhOH and thymol, slight soly. in H_2O with resultant slow resorption and good compatibility. W. O. E.

Content of mydriatic alkaloids of sun- and shade-dried Solanaceae drugs (Atropa, Hyoscyamus and Datura species). FLEMER KÖRNER. *Pharm. Zentralhalle* 72, 113-5 (1931).—Expts. show that the alkaloidal content of *Belladonna* and of *Hyoscyamus* is not materially affected by the manner of drying, whether in the sun or shade; with *Datura*, however, the amt. of alkaloids in shade dried material is about 10% greater than in that dried in the sun. W. O. E.

Scientific pharmacy. XVII. Production of organotherapeutic preparations. RAPP. *Pharm. Ztg.* 76, 260-1 (1931). cf. *C. A.* 25, 390.—A discussion of the various glands, their importance in pharmacy and methods of prepn. and evaluation. W. O. E.

Alcohol number of the D.A.-B.VI. F. GRAY. *Pharm. Ztg.* 76, 261-2 (1931).—A critical examn. of the official Ger. method showing that correct results are obtainable only by a thermometric check of the distn., and not by distg. off as prescribed, 13 and 11 cc. W. O. E.

Evaluation of flores cinse D.A.-B.VI. W. BRANDRUP. *Pharm. Ztg.* 76, 262 (1931).—The official provisional test prescribed for this drug is often inadequate in its evaluation. The following is suggested: on about 2 g. of the powd. sample in a filter pour about 5 cc. of CHCl_3 and evap. the filtrate to dryness in a porcelain dish. The residue, moistened with ale. KOH , should yield an orange to red color. W. O. E.

Extraction of ergot by the methods of D.A.-B VI and U. S. P. X, with notes on ammoniated tincture of ergot, B. P. FRANK WOKES AND G. K. FLINICK. *Quart. J. Pharm. Pharmacol.* 3, 594-625(1930).—A study has been made of the methods of prep. liquid ext. of ergot presented in the Ger and U. S. pharmacopoeias. The Ger method is on the whole less efficient than the Am. method but the difference is much smaller than at first seen from the work of Linnell and Randle. In the percolation of ergot with neutral 90% alc. as in the Ger process, the ext. efficiency is greatly affected by the degree of acidity of the ergot, because of the phosphate and other buffering substances contained therein. With the more acid ergots (pH below 5.5) neutral alc. may ext. the ergotoxine almost completely as it is extd. by acidified alc. (contg. 2% concd. HCl or 1% tartaric acid). With the less acid ergots (pH above 6.0) neutral alc. may ext. less than 0.5 the amt. of ergotoxine taken out by the acidified alc. Ammoniated alc., as in ammoniated tincture of ergot B. P., gives similar results to neutral alc. Whatever process of extn. is used the buffering substances exert considerable influence on the pH of ext. of ergot, both during extn. and in the completed product. If this pH is to be maintained at the optimum point for extn. and stability, as suggested by American workers, it will not be sufficient to employ a fixed proportion of acid, but control will be necessary by means of pH detns. The delatting of ergot increases the efficiency of extn. of both neutral and acidified alc. Complete removal of fat, without loss of ergotoxine, is very difficult to secure, and so-called "defatted" ergot may still contain a considerable proportion of fat. The concn. of dil. exts. of ergot should be carried out *in vacuo* below 37° with min. exposure to air. If these precautions are taken, concn. may be continued down to a soft ext. with the loss of less than 0.5 the activity, and it should be possible to obtain from an a. sample of Spanish or Portuguese ergot a soft ext. contg. from 0.5 to 1.0% ergotoxine. The variation in potency of concn. soft exts. suggests that sufficient care is not taken in their manif. In the Ger pharmacopoeial process for concg. the dil. percolate, a considerable amt. of activity may be lost in the pptn. and neutralization with Na_2CO_3 , and the U. S. P. method of concg. down to a soft ext. appears to give better results. W. O. E.

Biological standardization of tincture of aconite, B. P. F. J. DYER. *Quart. J. Pharm. Pharmacol.* 3, 626-66(1930).—Six tinctures of aconite were bought directly from different mfg. houses and examd. for their toxicity on rats and mice. Detns. were also made of the amt. of Et_2O -sol. alkaloid, and of total solids present in each. The examn. on rats revealed only quant. differences between the different tinctures, while that on mice was complicated by the emergence of qual. differences as well. The toxicities of the different tinctures both on rats and mice varied very widely, and the relative values as detd. on rats were found to be very similar to the relative amts. of Et_2O -sol. alkaloids present. *Conclusions*.—There is no uniformity in the toxicity of different samples of tincture of aconite sold as "B. P." at the present time. The figures obtained for the % of Et_2O -sol. alkaloids were either correct or they were incorrect. If correct, then there is no need for a toxicity test of tincture of aconite, because a detn. of the Et_2O -sol. alkaloids gives a parallel result. If, however, these figures were correct, then at least 4 of the 6 tinctures purchased did not comply with the B. P. If not correct, then it may be supposed that all the tinctures complied with the B. P., and the wide variation in toxicity of the different tinctures means that a toxicity test is required. W. O. E.

(Tobacco) smoking, nicotine limits and the determination of nicotine by the method of Pyl and Schmitt. FR. BOLM. *Z. Untersuch. Lebensmittel* 59, 602-7(1930).—Aside from minor modifications, the method of Pyl and Schmitt (*C. A.* 22, 1214) was found entirely satisfactory, and the observations on the influence of rate, intermittency, etc., of the smoking on the nicotine content of the smoke, were confirmed. Several samples of nicotine-free tobacco and cigarettes were examd. for nicotine content. None of the methods used in fixing the nicotine in the tobacco gave a nicotine-free smoke, hence the detn. of total nicotine in the tobacco is in itself sufficient in judging the tobacco. C. R. FELLERS

Ethereal oil from *Podocarpus nathorstii*, Don. I. KITSUJI NISHIDA AND HIDETAKE UOTA. *J. Agr. Chem. Soc. Japan* 6, 1078-86(1930). *Bull. Agr. Chem. Soc. Japan* 6, 82-3(1930).—By the distn. of the leaves 0.049% of ethereal oil was obtained α -pinene, camphene, β -pinene and cadinene were proved to exist in the vacuum distn. fractions. Caryophyllene and silyvestrene were not detected. Y. KIHARA

A method for the germicidal assay of soaps. H. C. HAMILTON AND FRED THISTLETHWAITE. *J. Lab. Clin. Med.* 16, 391-6(1931).—A method is described in which the skin of guinea pigs is used as a field for the testing of disinfectants commonly used on the skin. $EtOH$ applied before making injections is ineffective in any concn. A 0.5%

soln. of PhOH or tincture of I. in a 33% soln is effective Toilet or washing soaps are wholly ineffective A soap base in which HgI₂ is properly incorporated in a concn of 1% is effective

E. R. MAIN
Pyrethrum. ED ZACHAREWICZ *Prog Agr sci* 94, 569-70(1930)—Directions are given for cultivating the plant. The application of superphosphate is important as it increases the content of the plant's active principle, *oleoresin*

F W MARSH
The antiseptic power of uroselectan. G. DELEPIANE *Boll soc. ital biol sper.* 5, 822-4(1930)—Although uroselectan has feeble antiseptic power *in vitro* in the concn used clinically (4-6% in pyclographic tests), it definitely inhibits the growth of *B. coli* in the urine

PETER MASUCCI
The determination of some medical products by mercurimetry. AL IONESCO-MATIU AND A. POPESCU *Bull sci pharmacol* 38, 71-6(1931), cf *C A* 24, 5428—From a soln of about 1% *sparteine* mix 1 to 4 cc. with 5 cc. of the reagent of Mayer-Valzer Centrifugalize and wash the ppt. with 1% H₂SO₄. Dissolve in 5 cc. sulfonitric reagent, applying moderate heat, add 5 cc. more and 6 cc. H₂O, boil and, after cooling, destroy the nitrous vapors with KMnO₄. Ppt the lig with Na nitroprussiate (10% soln) and add 0.1 N NaCl until the turbidity disappears One cc. NaCl soln is equiv to 0.01423 g. of sparteine sulfate For the detn of *norocaine*, 1-5 cc. of a 1% soln is taken The method is the same as described, but 25 cc. H₂SO₄-HNO₃ is used The factor per cc. NaCl soln is 0.02 The same method is used for stavaine-HCl (factor 0.0155) and for *plasmochine* soln (factor 0.00926) With *plasmochine* tablets, 5 are dissolved in 10 cc. H₂O and 10 drops of H₂SO₄, filtered, brought to the vol. of 25 cc. and 5 cc. is employed for the detn In *plasmochine* dragées, the substance is detd together with quinine, the factor for the latter being 0.0066

A. E. MEYER
Some new color reactions of adrenaline. MARCEL PAGET *Bull sci pharmacol* 38, 77-8(1931), cf *C A* 25, 532—NaBrO reagent, diln 1/100, produces in a soln. of adrenaline of a concn higher than 1/1000 a color between a pale violet to intense red A violet coloration is still observed to a diln of 1/100,000 Adrenalone gives a yellow color The classification of adrenaline reactions in correspondence with the mol groups involved is shortly discussed

A. E. MEYER
A new method of determination of cocaine and its salts. JUAN A. SANCHEZ. *Semana med* (Buenos Aires) 1931, I, 487-9—After sapon of the cocaine, the PhCO₂H is detd Boil 0.1 g cocaine-HCl with 5 cc neutral EtOH and 10 cc 0.1 N NaOH for 15 min After cooling add 10 cc H₂O and titrate with 0.1 N H₂SO₄, using phenolphthalein as indicator The factor is 0.03755 when cocaine base is used and half that much when its salts are analyzed

A. E. MEYER
Studies in tobacco chemistry. IV. Methods of tobacco smoke analysis. N. J. GAVRILOV AND A. W. KOPERINA *Biochem Z.* 231, 23-32(1931), cf *C A* 24, 882—An app. is described for absorbing all the products formed in glowing tobacco. The analyses of total N, nicotine N and NH₃ N vary only 3-4%

S. MORCULIS
Experiments on *Cymbopogon nardus* Rendle and on *Cymbopogon martin* Stapf, var. *Sofia*. FRANCESCO BRUNO *Parfums de France* 9, 8-14(1931)—*C. nardus*, planted in Italy (Palermo) on March 15, 1926, had grown to an av height of 80 cm. on Aug. 1, it was cut and distd on Aug. 1 and on Oct. 1, yielding 14.6 and 14.2 tons of leaves (total 28.8) per ha., which gave 0.5% and 0.4% of oil, equiv. to 73 kg and 56.5 kg, or a total of 129.5 kg. per ha for the 2 cuttings The corresponding yields for 1927 (cuttings on July 26 and Oct. 12) were: 28.3 and 17.1 tons, 0.535 and 0.362%, 151.4 and 61.8 kg per ha. Analysis of this oil in the Schummel lab. gave: d₄ 0.8933, α (100 mm tube) -2°11', n_D 1.4720, acid no. 0.9, ester no. 58.8, Ae no. 259-5 (= 88.6% esters), sol in 1 or more vol of 80% alc. Analysis of the oil distd in 1928 gave: d₄ 0.8926, α +1°, total geraniol 84.67%, aldehydes (citral + citronellal by Schummel's phenylhydrazine method) 44.10%, free geraniol 30.62%, other alcs 9.95%, citronellal (Dupont and Labaune method) 33.10%, citral and other aldehydes (by difference) 11%, sol in 1.3 vol. of 80% alc and in 1.4 vol of 70% alc. Leaves harvested during the cold season (Dec. 18) yielded 0.25% oil (as compared with 0.535% from the July harvest) which gave the following results on analysis: d₄ 0.872, α +1°, total geraniol 70.36%, free geraniol 32.6%, aldehydes (citral + citronellal by Schummel's phenylhydrazine method) 52.5%, citronellal (Dupont and Labaune) 27.34%, citral and other aldehydes (by difference) 24.5%, sol in 1 vol. of 80% alc. and in 1.2 vol of 70% alc. A portion of this same harvesting was allowed to dry till it had lost 72% in wt. The dried leaves gave the same yield of oil as the fresh leaves. The drying of the leaves had the following effects on the compn. of the oil (1) Total geraniol remained at 70.36%. (2) Free geraniol fell from 32.4% to 2.4%. (3) Citronellal increased from 27.34% to 44.79%. (4) Citral and other aldehydes fell from 24.5% to 10.21% A further har-

N_3As_2 which is insol in water but sol in MeOH , EtOH and acetone is obtained by a process described in detail and is suitable for combating syphilis and sleeping sickness.

Iodinated, acetylated condensation products of the di(aminobenzyl) derivatives of methylenedisuccinic acid with terpin hydrate. SAMUEL L. SUMMERS. U. S. 1,793,028, Feb. 17. Products insol in water but sol in alc. and which may be used as antiseptic, antineuralgic, antitubercular and antiscarlatinal remedies are described with details of manuf.

Neutral calcium salt of inositolhexaphosphoric acid. "CRISTALLO" A-G. Swiss 141,522, Dec. 8, 1928. Vegetable matter rich in inositolhexaphosphoric acid and its salts are extd. with water and the neutral Ca salt pptd. from the ext. by addn. of CaCO_3 . The Ca salt is filtered off and used in therapy. The remaining liquid is rich in vitamins which are unaffected by the presence of the carbonates, and can also be used for medicinal purposes.

Granular anhydrous citric acid. RICHARD PASTERNAK and FRANKLIN G. BREHMER (to Charles Pfizer & Co.). U. S. 1,792,657, Feb. 17. Citric acid is dehydrated by drying at temps. of 35-50°, in order to obtain a product suitable for tablets, effervescent salt compns., etc.

Isolation of scillarene A from squills. CHEM. FAB. NORM. SANDOZ. Swiss 142,364, Feb. 19, 1929. Addn. to 134,217 (C. A. 24, 1706). The cryst. almost H_2O -insol. ext. from squills, which has a strong action on the heart and contains glucoside and tannoid, is extd. with water and the insol. part added to a mist. of water and org. solvent. This is treated with a tannin pptg. agent to remove the tannin and scillarene A is obtained from the filtrate.

Anti-syphilis preparation. SALO BERGEL. Ger. 502,655, Feb. 19, 1925. Virulent syphilis spirochetes are killed, dissolved in lecithin emulsion and the resulting prepn. injected into the pleural or abdominal cavities or in the ear veins of animals. The blood, abdominal and pleural exudations, lymphatic glands, milk, thyrus and ovaries are then extd. with glycerol and soda or dil. acid soln. to obtain the anti-syphilis prepn.

"Cancer remedy." H. LERTNER. Brit. 337,797, July 8, 1929. Gall bladders of animals are extd. with alc., ether or other solvent and the ext. is evapd. to dryness, and the resulting product is preferably dissolved in an aq. soln. of KCl or other suitable K compd. and used for injection into the affected part.

Vaccines. PHILADELPHES KIRBY. Ger. 517,204, Feb. 12, 1929. The usual bacteria are cultivated on nutrient mediums which are inoculated also with filtrates from cultures of the known amoeboid forms. The coocoid growths so obtained are killed or enfeebled by the action of heat or disinfectants.

Treating blood, bandages, etc., to prevent bacterial decomposition. I. G. FAKSBERG and A. G. Brit. 337,473, May 1, 1929. The materials are treated with small addns. of substances such as methylenebis 2,4-dichlorophenol, the condensation products of formaldehyde with 2 mol. proportions of *p*- or *m*-chlorophenol, the condensation products from *p*-chlorobenzaldehyde or *p*-diethylaminobenzaldehyde and 2 mol. proportions of *p*-chlorophenol, a product obtained by treating the condensation product of cyclohexanone and 2 mol. proportions of phenol with a glacial HOAc soln. of Br or with a K bromate-bromide soln., or a product obtained by heating 2,6-dimethylol-4-chlorophenol with 4-chlorophenol in the presence of 50% H_2SO_4 .

Filter for essences, etc. MARINO FISANI. Swiss 142,143, July 29, 1929. The filter is particularly suitable for removing particles of water contained in the essence, etc.

Medicating tobacco with iodine. FRANK P. STRUCKLER. U. S. 1,793,723, Feb. 24. I is first converted into a fluid state, as by vaporization or dissolving in alc. and is caused to crystallize in contact with tobacco to render the latter suitable for use as a therapeutic source of I.

Tobacco smoke treatment. H. PÄFFGEN. Brit. 333,006, Nov. 19, 1929. Poisonous constituents of tobacco smoke such as nicotine, pyridine and "brenzole" are sep'd by passing the smoke through a dry absorption medium such as activated C or silica gel which has been previously sat'd. with a volatile substance of suitable character to prevent absorption of aroma imparting substances also (as by use of smoke from combustion of nicotine free tobacco waste, peppermint oil, NaOH , Na_2CO_3 , sucrose, NaCl , tannic or barbituric acid, O_2 or Na hyposulfite, or by merely leaving 20-25% of water in silica gel if this is used).

18—ACIDS, ALKALIES, SALTS AND SUNDRIES

E. M. SYMMES

The concentration of weak nitric acid. G. PAYRAS. *Rev. gén. mat. plastiques* 6, 147-51(1930).—A description is given of the Chem. Construction Co.'s Fr. pat. 656,108 (C. A. 23, 4026).

Vanadium as a catalyst for sulfuric acid manufacture in America. ENRIO ALHOTRO. *Acta Chemica Fennica* 3, 112-4(1930).—The characteristics of V catalysts now used in plant operation are given. The best results are obtained by the Selden V mass and the new converter developed by A. O. Jaeger, which give as high as 98% conversion. In present-day plant operation V gives a higher av. conversion than Pt, because Pt is contaminated with traces of As and Cl. The price of Pt is 184-253 times the price of V, but the high return value of Pt catalysts and the high price of licenses for V catalysts make this proportion lower. H₂SO₄ plants which were operated or built in 1929 and which use the Selden converter system have a total capacity of more than 300,000 tons of H₂SO₄ per year. The total production of H₂SO₄ by the contact process in the U. S. during the same year was 1,570,000 tons as 100% acid.

S. A. KARJALA

Potash. GEO W. STOCKING. *Chem. Markets* 23, 247-52(1931).—The industry in Germany is described.

E. H.

A study of the properties of polyhalite pertaining to the extraction of potash. VI. Experiments on the production of potassium chloride by the evaporation of leach liquors from decomposition of uncalcined polyhalite by boiling saturated sodium chloride solutions. H. H. STORCH AND F. FRANK. *Bur. Mines, Rept. of Investigations* 3062, 7 pp(1931), of C. A. 25, 1338.—Evapn. of 90% of the H₂O, interspersed with 3 crystn. steps, will yield 78% of the K₂O as crude KCl, the bulk of the impurities being NaCl. This product may be refined readily by recrystn. to produce pure KCl. A preliminary estimate indicates about \$20 a ton as the cost of production at the plant. A. H. E.

The acids and acid salts which contaminate cream of tartar. FILIPPO PERCIVALPOSCO. *Atti III congresso naz. chim. pura applicata* 1930, 520-5.—Some cream of tartar contains appreciable quantities of Ca(C₄H₄O₆)₂ instead of KCl.

E. M. S.

Quinquennial review of the mineral production of India for the years 1924 to 1928. Phosphates. E. H. PASCOE. *Records Geol. Survey India* 64, 413-21(1930).—Apatite is known in many parts of India, some deposits assaying 20-25% P₂O₅. Nodular deposits of lime phosphate also occur.

ALDEN H. EMERY

A brief history of the world's phosphate rock production. A. N. GRAY. *Superphosphate* 4, 25-46(1931).—Historical notes relating to the discovery of phosphate deposits and the early production of phosphate rock in all parts of the world are given. The world's annual production of phosphate rock by countries is given for the years 1847 to 1929.

K. D. JACOB

Dehydration of salts of phosphoric acid. S. S. DRAGUNOV. *Udobrenie i Urazhenie (Fertilizers and Crops)* 2, 409-16(1930).—Graphs show the transformations of the salts of Na, K and Ca or phosphoric acid upon drying. From these it is possible to judge the % of pyro- and meta-modifications formed, as well as the speed of the transformations at various temps. With the loss of H₂O of crystn. CaHPO₄·2H₂O becomes less sol. in citrate soln. At 100° only 4% of the H₂O of crystn. was lost in 1.5 hrs. At 150° 19% was lost in 20 min. It is important not to dry phosphate very much above 100° for any period of time (not more than 30-50 min. at 100°). Ca metaphosphate, insol. in acid, can be hydrated in the autoclave at high pressures and brought back to the sol. state.

J. S. JOFFE

Utilization of the gases obtained in the process of volatilizing phosphorus and the production of soluble phosphates. A. P. DRYAKOV. *Udobrenie i Urazhenie (Fertilizers and Crops)* 2, 397-409(1930).—The relation of H₂O and O from the air to the system; P, PH₃, CO, H, in which a reaction of selective oxidation of generator gas takes place was investigated. A modification of the app. used by Britzke and Pestov (cf. C. A. 24, 218) was necessary. A description and drawing of the new app. are given. The exptl. data obtained show that the oxidation of P with H₂O over CaO at a temp. from 550° to 950° is completed to phosphoric acid at fairly rapid velocity of the gas stream. Very little PH₃ is found in the gases, and the sum of CO + H is not decreased. The materials used for oxidation contained no phosphates. In the oxidation of the gases with air no excess of O was necessary in the process, except what was needed for the oxidation of P. At the lower limits of temp. no O is used up by the CO. The generator gases used in this manner can be used again for heating or other purposes. After

the process was worked out with lime, NaCl was tried. On account of the volatilization of the NaCl and even the Na phosphates formed, it was necessary to keep the temp between 400° and 500°. With H_2O no oxidation of P took place, but with O it did, and the generator gases were free from P, but slightly diluted with N. J S J.

Calcium arsenate. Economic data regarding its development in Mexico. PABLO HORN Y HORN. *Rev quim* 6, No 2, 11-9(1930); cf C. A. 25, 1039— As_2O_3 can be obtained from smelters in Mexico, but is unprofitable. If made into $Ca_3(AsO_4)_2$, 2,200,000 kg. could be used on cotton fields in Mexico. E. M. SYMONS

Magnesium compounds (other than magnesite). PAUL M TYLER. *Bur Mines, Information Circ* 6406, 19 pp (1931). cf C. A. 25, 561—Properties, uses, methods of production, domestic and foreign production, imports, exports, prices and manufacturers of $MgCO_3$, $MgSO_4$, calcined magnesia and $MgCl_2$ are discussed. A H E.

Separation of the constituents of sylvinite in the form of carbonates. FIDOLARD URBAIN. *Compt rend* 192, 232-3(1931)—Almost complete pptn of K in the form of acid carbonate of K and Mg may be obtained, and the Cl of KCl may be removed in the form of NaCl, by using $NaHCO_3$ according to the reaction, $MgCO_3 + NaHCO_3 + KCl = MgCO_3KHCO_3 + NaCl$. This reaction may be applied to complete sepn. of K and Na in the mixt. of chlorides known as sylvinite, which is a most important mineral in potash mines. After sepn. of the solid phase the soln. is treated to recover $NaHCO_3$. The cycle of operations permits quant. recovery of K_2CO_3 and Na_2CO_3 from sylvinite. ALICE W. EFFERSON

The economic significance of zirconium and its compounds. J F CORRIGAN. *Ind Chemist* 7, 97-100(1931). E H

Extraction of beryllium, cesium and rubidium from beryl. C JAMES, H C. FOGG AND E D COUGHLIN. *Ind Eng Chem* 23, 318-20(1931)—The finely ground mineral is fused with CaO , and the resulting slag is ground, decompd. with H_2SO_4 , and heated to dehydrate the SiO_2 . The mass is stirred with hot water and the filtrate is concd. to ppt. out K, Cs and Rb alums. The remaining $Al_2(SO_4)_3$ is removed as NH_4 alum, and the Fe, after oxidation, is pptd. with dil. NH_4OH , the last traces being removed by H_2S under slight pressure. The Be is then pptd. as the basic carbonate by the addn. of $(NH_4)_2CO_3$. The method can be adapted to the extn. of Be from other silicate minerals. K. D JACOB

Activated carbons. G SIMON. *Caoutchouc & gutta-percha* 26, 14,357-9, 14,468-70, 14,516-9, 14,551-3, 14,594-5, 14,632-6, 14,701-2, 14,742-4, 14,788, 14,821(1929); 27, 14,866-8, 14,903-4, 14,940-1, 14,983-5, 15,022-3, 15,064-5(1930)—After the important phys. properties of activated carbons are described it is shown that the activity does not depend to any appreciable extent upon the nature of the gas from which the vapor is removed. If care is taken to avoid too high temps., it is possible to recover $MeOH$, $EtOH$, $AcMe$, Et_2O , $EtOAc$ and CS_2 unaltered from activated carbons in which they have been adsorbed from air. The most rapid method for expelling the adsorbed vapors from carbons, without chem. alteration, is by treatment with steam. Below 130° the steam may be adsorbed to some extent, but above this temp. it is possible practically to avoid such adsorption. Investigations of the tarry matter retained by the carbons after removal of C_4H_6 from coal gas showed that surprisingly little real tar was present, probably because of rapid resinification of the major portion of such small quantities as may be present in the gas, by contact with the charcoal. Inactivation of carbons in the removal of C_4H_6 by coal gas is attributed to the effect of S, which converts the active C into CS_2 . Reactivation by calcination before and after treatment with HCl is explained tentatively by the decompn. of the CS_2 by Fe naturally present. The acid not only decomposes the FeS formed in the first calcination but brings about a uniform and intimate dispersion of the regenerated Fe salt in readiness for the second calcination. Similarly, it is possible to remove the inactivating S and so to revivify the C by impregnation with $FeCl_3$ soln., followed by heating, e. g., for 2 hrs. at 300°. B C. A.

A western nitrogen-fixation plant. C N WESTON. *Ind Eng Chem*, 22, 1093-1103(1930)—A description is given of the unique plant of the American Nitrogen Products Co., which was destroyed by fire in 1927. Nitrites were manufactured by the arc process and the details of operation are described. The final product contained 96.5-99.0% $NaNO_2$. J. R. ADAMS

The synthesis of ammonia from its elements in electric discharges. M. ALSFELD AND E. WILHELMY. *Ann. Physik* 8, 83-123(1931)—The yield of NH_3 obtained by passing a 3 l. H_2 , N_2 gas through a silent discharge in a Siemens tube has been detd. at temps. of 25° to 300°. The rate of reaction increases slowly with the temp. up to 130° and rapidly from there to 300°. The sudden change in temp. coeff. at about the crit.

temp of NH_3 is tentatively explained as due to a sudden decrease in the no. of NH_3 mols adsorbed at such points on the surface of the Siemens tube as to inhibit NH_3 formation. Townsend, corona, glow and arc discharges were compared as regards their effectiveness in synthesizing NH_3 in a stream of $\text{H}_2\text{-N}_2$ gas passed through them. NH_3 was detd in all cases by titration. In the Townsend discharge no NH_3 was obtained. In the neg corona the yield of NH_3 was const from 300 to 780 mm. pressure at a given c d in the glow discharge at pressures below 300 mm it increased linearly with decreasing pressure. At 760-mm pressure, NH_3 formation in a corona discharge increased linearly with the current until an arc was formed; it then decreased markedly. In the glow discharge the most rapid rate of synthesis occurred in the cathode region, the pos column was less active. Fe, Cu and W as cathode materials all produced about the same yield of NH_3 , Pt cathodes were 30% more effective, possibly because of some synthesis occurring on the Pt between H atoms and N_2 mols. In the neg glow, the yield of NH_3 increased linearly with the current and for a given c d increased with the rate of gas passage. It decreased slightly as the cathode temp was increased. In the anomalous cathode fall, the yield of NH_3 per unit current was greater than in the normal cathode fall. The max rate of NH_3 synthesis was found in a gas mixt. contg 60% N_2 and 40% H_2 . At the same compn a min in the cathode fall was noted.

P. H. EMMETT

The use of pyrethrin against bed bugs. ALBERT GUILLAUME. *Bull. sci. pharmacol* 38, 80-4 (1931).—A concd soln of pyrethrins in MeOH dild. with 25 parts of H_2O proved to be a very effective spray against bed bugs. The eggs are not destroyed. For the extinction, the buildings must be treated 3 times a year in March-April, in July and in August-September.

A. E. MEYER

Utilization of the aqueous extracts from olives [recovery of KCl] (Crusa) 27. NaHCO_3 (Brit. pat. 338,007) 15. Apparatus for catalysis [in oxidation of NH_3] (Swiss pat. 141,303) 1.

Nitric and sulfuric acids. I. G. FARBER AND A. G. Brit. 337,406, June 13, 1929. In simultaneous production of HNO_3 and H_2SO_4 by scrubbing gas mixts. contg N and S oxides with sulfuric or nitrosylsulfuric acid, as described in Brit. 301,232 (C. A. 23, 4026), a sufficient proportion of N oxides is present that all the SO_2 is converted into H_2SO_4 or nitrosylsulfuric acid or a mixt. of H_2SO_4 and HNO_3 . Various details of procedure are described.

Sulfurous acid. SOC. ANON. POUR L'IND. CHIM. A. BALE. Swiss 142,731, May 15, 1929. Conc'd H_2SO_4 is prepd from gases contg SO_2 by absorbing the SO_2 by water and aromatic amines, driving off the SO_2 by heating and reabsorbing in water.

Phosphoric acid. KUNSDÜNGER-PATENT-VERWERTUNGS A.-G. Swiss 141,866, Apr 3, 1929. H_3PO_4 soln is obtained by causing mineral acid to react with crude phosphates and pptg the Ca of the crude phosphate by sulfate ions. Examples are given. Cf. C. A. 24, 471.

Ammonia synthesis. LOUIS C. JONES (to Chemical Engineering Corp.). U. S. 1,794,903, March 3. The mixt. of H and N for NH_3 synthesis is purified by circulating the gases under pressure in contact with a catalyst and introducing CO_2 into the circulating system after compression and before catalysis, to react with gases and convert impurities in the system into solid products, which latter with CO_2 are simultaneously removed (suitably as $(\text{NH}_4)_2\text{CO}_3$ deposited on Raschig rings) while maintaining a continuous circulation of the gases under pressure through the system as the gases pass to the catalyst. Cf. C. A. 24, 4123.

Ammonia synthesis. HERBERT A. HUMPHREY (to Atmospheric Nitrogen Corp.). U. S. 1,794,231, Feb. 24. Solid carbonaceous fuel is burned in the presence of steam, the gaseous products are subjected to catalytic action to cause reaction of CO with H_2O to form a volume of H substantially equal to that of the CO , a proportion of 3:1 between the H and N is established, O compds. are removed from the gas mixt. and the latter is catalyzed to convert part of the gases into NH_3 , the NH_3 is sepd. and unconverted gas is returned to the NH_3 -forming catalyst through a closed, periodically bled, circulatory system. The air and steam used with the fuel in the first step of the process are highly preheated and a temp. of about 1300° is maintained in the gasification zone. The resulting gaseous products are withdrawn from the burning fuel and heat from them is transferred to the entering air and steam while contact of the gaseous products with colder solid fuel is prevented. A substantially CH_4 -free gas is thus produced which lengthens the periods between the bleedings required.

Ammonia-air mixture. HARRY PAULING. Ger. 491,961, Oct. 16, 1926. A homo-

gencous NH_3 -air mixt. is obtained by absorbing the NH_3 in a circulating water current and then blowing air through the soln. App is described Cf C. A. 25, 1041.

Catalytic apparatus suitable for ammonia oxidation. C. TONIOLO and "AZOCENO" (SOCIETÀ ANON PER LA FABBRICAZIONE DELL'AMMONIACA SINTETICA E PRODOTTI DERIVATI) Brit. 337,680, July 25, 1929. In app comprising a plurality of superposed layers of Pt or other suitable metallic gauze, the layers first impinged upon by the reaction gases are of coarser mesh than the subsequent layers which are less subject to attack by the gases. Various details of construction are given.

Working up mixtures containing carbon dioxide and ammonia. I G FARSEVINO A-G (Ernst Jancke and Frich Rablitz, inventors) Ger 514,893, Nov 6, 1929. Mixts. contg CO_2 and NH_3 or their compds., especially the mixt. resulting from the prepn. of urea from $\text{NH}_3\text{COONH}_4$, are prepared by treating the mixt. with liquid NH_3 . Thus, a mixt. of urea, $\text{NH}_3\text{COONH}_4$, and water is washed with liquid NH_3 at -36° , giving a soln. of urea and water in NH_3 and a pure $\text{NH}_3\text{COONH}_4$ residue. The urea is then recovered by evapn.

Apparatus for production of ammonium salts. "MONTECATINI" SOCIETÀ CEMERALE PER L'INDUSTRIA MINERARIA ED AGRICOLA. Swiss 142,051, June 7, 1929. See Brit. 313,446 (C. A. 24, 1186).

Cyanamides. N. CARO and A. R. FRANK. Brit. 338,023, Dec. 6, 1928. In prepg. cyanamides of Ca or Mg or both by reaction of the corresponding carbonates and NH_3 as described in Brit. 279,811 (C. A. 22, 3024), the removal of the water produced in the reaction is accelerated by substances such as C, carbonaceous materials or P_2O_5 which are preferably mixed or briquetted with the starting materials. When C or carbonaceous materials are thus used, oxides or compds. forming oxides on heating may be substituted for the carbonates, and when P_2O_5 is used a composite fertilizer is obtained. Cf. C. A. 25, 1041.

Phosphates and hydrogen. BAYERISCHE STICKSTOFF WERKE A-G (Vladimir N. Ipat'ev, Nikodem Caro and Albert R. Frank, inventors) Ger 514,830, Mar 27, 1928. See Brit. 308,784 (C. A. 24, 472).

Nitrogen oxides, etc. N. CARO and A. R. FRANK. Brit. 337,847, Aug 13, 1929. N_2O_4 or HNO_3 in concd. form is obtained from mixts. contg N oxides, O and water vapor (such as products of NH_3 oxidation) by a 3-phase process in which part only of the water is first removed without much oxidation, residual water is sep'd. after or during oxidation (which is promoted by lowering the temp.) as a tetroxide contg HNO_3 of low water vapor pressure, while in the third phase the gas may be cooled or treated with substances which lower the vapor pressure of the tetroxide or have absorbing properties. Numerous details of procedure are given.

Precipitating metals as sulfides. NIELS C. CHRISTENSEN U S 1,793,906, Feb. 24. Solns. of metals, the sulfides of which are sol. in such solns. if slightly acid, e. g., solns. contg Zn, Fe and Mn, are treated simultaneously with CaCO_3 and H_2S to ppt. the metal sulfides and form CO_2 and the latter is used for treating a $\text{Ca}(\text{SH})_2$ soln. to form H_2S and CaCO_3 for further use in the process to ppt. addnl. metal.

Purifying zinc salt solutions. I G FARSEVINO A-G. Brit. 337,822, Aug 8, 1929. Solns. obtained from chlorinated burnt pyrites are freed from Fe by oxidizing the latter with an alk. earth chlorate such as $\text{Ca}(\text{ClO}_3)_2$ and then pptg. with quicklime. Cu is first removed by addn. of Fe.

Ammonium sulfate. UNION CHIMIQUE BELGE, SOC. ANON. Brit. 338,373, May 17, 1929. In prepg. $(\text{NH}_4)_2\text{SO}_4$ as described in Brit. 307,037 (C. A. 23, 5281), with pptn. of the sulfate as crystals by addn. of NH_3 , the pptn. is localized in a part of the app. in which the ammoniacal soln. is mixed with a large quantity of cooler satd. soln., and a portion of this soln. is continuously circulated through a cooler so as to withdraw the heat caused by the soln. of the NH_3 . Various details of app. are described. Cf. C. A. 24, 715.

Decomposing ammonium sulfate. WILHELM BEBELSMANN and ALFRED C. BECKER. Ger 517,495, Apr 8, 1928. The thermal decompn. of $(\text{NH}_4)_2\text{SO}_4$ to NH_3 and NH_4HSO_4 is effected in vessels made of SiO_2 . Shallow vessels are preferred. The temp. should be maintained above the m. p. of the sulfate melt. Addnl. details are given.

Barium aluminate. KALI CHEMIE A-G (Fritz Rothe and Hans Brenck, inventors) Ger 514,891, July 31, 1926. Alumina-rich material is heated with BaSO_4 to a temp. of at least 1000° in a current of steam and oxidizing or indifferent gas. Heavy spar may be used as the starting material in such amt. that besides the $\text{Ba}_3(\text{AlO}_4)_2$ formed, there is also produced silicates of the types BaSiO_3 or Ba_2SiO_4 or the corresponding Ti compds. In an example, 100 parts of heavy spar contg 92.63% BaSO_4 and

3.72% SiO_2 are mixed with 56 parts bauxite contg. 52.86% Al_2O_3 , 2.71% SiO_2 and 2.84% TiO_2 . The mixt. is then heated to 1200° in the presence of steam and air, producing $\text{Ba}_2(\text{AlO}_4)_3$ and BaSiO_3 . Other examples are given.

Barium carbide. INTERNATIONAL INDUSTRIAL & CHEMICAL CO., LTD. Brit. 338,003, Nov. 21, 1928. See Fr. 682,479 (C. A. 24, 4362).

Calcium nitrate. APPAREILS ET EVAPORATEURS KESTNER. Ger. 514,589, Mar. 21, 1928. In the prepn. of $\text{Ca}(\text{NO}_3)_2$ from CaCO_3 and HNO_3 , the CaCO_3 is loaded into a tower mixed with granular substance not acted on by the HNO_3 , to ensure a passage for the acid through the mass. App. is described. Cf. C. A. 24, 5442.

Calcium oxide. CONSORTIUM FÜR ELEKTROCHEMISCHE INDUSTRIE G. M. B. H. Ger. 514,715, July 19, 1927. CaO prepd. by heating $\text{Ca}(\text{OH})_2$ to $600\text{--}800^\circ$ in a gas current, can be used as a dehydrating agent for liquids or molten substances by mixing therewith. Thus, 95% alc., if heated with CaO prepd. as above, reaches a purity of 99.6% in about 15 min. Other examples describe the dehydration of Na alcoholites by CaO as thus prepd.

Siliceous calcium oxide compounds. FRIEDRICH PISTOR. Ger. 514,742, Aug. 7, 1927. Sol. dry compds. are obtained by adding CaO to concd. solns. of siliceous substances. Thus, burned lime is added to Na or K water glass. The products are used in the cold cement, drying and artificial mass industries.

Dicalcium phosphate. ERICH THULO and ADOLF HEICKE. Ger. 517,181, Feb. 12, 1928. See Fr. 677,191 (C. A. 24, 3091).

Solid carbon dioxide. AKT. GER. FÜR MASCHINENFABRIKEN ESCHER, WYSS & CIE, SWISS 141,893, Oct. 19, 1929. Liquid CO_2 under pressure is allowed to expand in stages. Cf. C. A. 24, 4362, 4363.

Solid carbon dioxide. MIDDEN-EUROPEISCHE OCEETOOMMAATSCHAPPIJ. Ger. 514,717, Sept. 4, 1928. App. for suddenly allowing liquid CO_2 under pressure to expand to form CO_2 snow, and for allowing further reduction of pressure to form CO_2 ice is described. Cf. C. A. 24, 1474.

Carbon dioxide from limestone. P. BROTHERHOOD, LTD., H. M. DUNKERLEY and CARBON DIOXIDE CO., LTD. Brit. 338,369, Nov. 19, 1929. Limestone from a crusher is placed in retorts fed with steam from a boiler which is heated by waste heat from the retort furnace, and the mixt. of steam and CO_2 from the retort is used to work a prime mover, exhaust steam is condensed, and the CO_2 is collected. App. is described.

Carbon disulfide. ERICHARD LEGELER and HERMANN KOB (to I. G. Farbenind. A. G.). U. S. 1,793,181, Feb. 17. Various details are described of app. comprising an elongated main retort adapted to contain a charge of C, at least one elongated supplementary retort positioned along and in contact with the main retort in the same heating chamber, and a device for introducing liquid S at one end of the supplementary retort and for leading superheated S vapors from the opposite end of this retort into the main retort.

Carbon disulfide. ERICHARD LEGELER and EUGEN FISCHER (to I. G. Farbenind. A. G.). U. S. 1,795,132, March 3. See Brit. 333,090 (C. A. 25, 564).

Carbon disulfide. CHEM. FAB. KALK G. M. B. H. and HERMANN OERDM. Ger. 517,337, July 9, 1929. CS_2 is prepd. by leading a mixt. of H_2S and CO_2 with or without inert gases such as N, over glowing C. The initial gas may be obtained from coal distn. gases by washing with water. Cf. C. A. 24, 2333.

Iron-free chromic chloride from ferrochromium. I. G. FARBEIND. A. G. (Johannes Brode and Carl Wurster, inventors). Ger. 514,571, Sept. 11, 1928. Granulated Fe-Cr is heated to $300\text{--}600^\circ$ in a heat-insulated container in a current of Cl and the temp. maintained at such a value that the FeCl_3 formed sublimes, leaving CrCl_3 . An example is given.

Magnesium carbonate; ammonium sulfate and chloride. KALI CHEMIE A. G. Ger. 517,496, Oct. 17, 1928. The manuf. of $\text{MgCO}_3 \cdot 3\text{H}_2\text{O}$ and $(\text{NH}_4)_2\text{SO}_4$ by treating a soln. of MgSO_4 with $(\text{NH}_4)_2\text{CO}_3$ is effected in two stages, the first stage involving the reaction $\text{MgSO}_4 + 2(\text{NH}_4)_2\text{CO}_3 \rightarrow (\text{NH}_4)_2\text{SO}_4 + (\text{NH}_4)_2\text{CO}_3 \cdot \text{MgCO}_3 \cdot 4\text{H}_2\text{O}$. The double salt is filtered off and treated with a soln. of MgSO_4 to effect the reaction $(\text{NH}_4)_2\text{CO}_3 \cdot \text{MgCO}_3 \cdot 4\text{H}_2\text{O} + \text{MgSO}_4 \rightarrow 2\text{MgCO}_3 \cdot 3\text{H}_2\text{O} + (\text{NH}_4)_2\text{SO}_4$. MgCl_2 may be treated similarly, with production of NH_4Cl . The products are purer than those obtained when the reaction is effected in one stage. An example is given.

Magnesium sulfate. EMILIO ROBOLO. Swiss 142,436, Mar. 16, 1929. MgSO_4 is recovered from the waste water of artificial silk manuf. by addn. of $\text{Ca}(\text{OH})_2$, which causes pptn. of CaSO_4 , $\text{Mg}(\text{OH})_2$ and $\text{Zn}(\text{OH})_2$. The ppt. is sepd., suspended in water, and treated with CO_2 to ppt. $\text{MgSO}_4 \cdot \text{Na}_2\text{SO}_4$, may also be recovered from the residual liquid.

Magnesium sulfate solution. KALI CHEMIE A-G Ger. 514,590, Nov. 30, 1929. Ifot highly concd $MgSO_4$ soln is prepd by forming an unsatd soln. of the salt from natural kieserite and satg with anhyd kieserite

Potassium, calcium and magnesium sulfates from materials such as polyhalite. EUGENE P SCHUCH U S 1,794,552, March 3 Ore such as polyhalite is dehydrated by heating K_2SO_4 and $MgSO_4$ are extd with hot water, undissolved $CaSO_4$ is sepd from the soln K_2SO_4 is crystd from the soln and sepd and the soln is further treated for recovery of $MgSO_4$ U S 1,794,553 relates to a similar process Numerous details of procedure are given

Potassium and magnesium sulfates from polyhalite. EUGENE P SCHUCH U. S. 1,794,551, March 3 After dehydrating polyhalite to remove its H_2O of crystn, the material is extd with hot water to produce solns contg over 5% of anhyd salts

Potassium nitrate FÉLIX JOURDAN SWISS 141,306, Mar 30, 1929 KNO_3 , with $Al(NO_3)_3$, $Fe(NO_3)_3$, etc., is obtained by treating leucite with a mixt. of N-O vapor, $FINO_3$ vapor and water vapor

Sodium carbonate. W MANN Brit 337,401, July 30, 1929 Finely cryst. soda is made by mixing, in an edge runner mill, finely ground Solvay soda with a suitable proportion of water

Sodium sulfide. WOLF J MÜLLER and FRIEDRICH KLEMA Ger. 514,954, May 15, 1928 See Austrian 116,564 (C. A. 24, 2844).

Titanium dioxide. SIMON J. LUBOWSKY. U. S. 1,793,501, Feb 24 A finely ground mixt of rutile and magnesia is heated to form a Mg titanate, followed by acidification with H_2SO_4 , chilling the soln to deposit $MgSO_4$, removal of the mother liquor and heating it to a high temp. to produce metatitanic acid by hydrolysis with regeneration of H_2SO_4 , and the metatitanic acid is washed and calcined.

Zinc carbonate and oxide. I G. FARBENIND A-G Brit 337,792, Aug 7, 1929 A soln. of a Zn salt such as $ZnCl_2$, of a concn less than 1.5 N, is added in slight excess to an alkali metal carbonate soln. of a max concn of 1.5 N at a temp. of 60-70°, the $ZnCO_3$ ppt formed is filtered off, washed, dried, ground to pass through a wire sieve of 350 mesh per linear in. and then calcined below redness (suitably at 350°).

Crystals of materials such as metals or salts. PERCY W BRIDGMAN U. S. 1,793,672, Feb 24 A seed crystal is placed adjacent molten substance and a portion of the seed crystal is melted and the surface of solidification is advanced from the seed crystal into the substance An app is described

Borax crystallization. ROBERT B PEET (to American Potash & Chemical Corp) U S 1,792,863, Feb 17. An emulsion colloid such as oleic acid is added to an aq. soln which ppta substantially only borax on cooling and the soln is cooled while being agitated, to ppt. borax The added colloid prevents aggregation of the pptd borax crystals

Treating crude sylvinitic potash salts containing kieserite. THEODOR THORSSSELL and AUGUST KRISTENSSON U S 1,794,259, Feb 24 The kieserite in the crude material is hydrated and the salts are treated with an ammoniacal soln. of NH_4Cl , the K_2SO_4 contg Na_2SO_4 is removed and the mother liquor is treated with CO_2 for pptn of Mg NH_4 carbonate The mother liquor remaining is cooled to remove a portion of the NH_4Cl present and the liquor still remaining is treated with CO_2 to ppt $NaHCO_3$ U S 1,794,260 describes a process for the production of K_2CO_3 from crude sylvinitic salts, with simultaneous production of Na_2CO_3 , by treating the crude salt with NH_3 , carbamate in substantially anhyd liqnd NH_3 , converting the resulting K and Na carbamates into the bicarbonates by treatment with water, sepg the bicarbonates from each other and from assoc impurities (suitably by known methods) and converting the bicarbonates into carbonates.

Concentrating phosphate-bearing material such as phosphate rock. WM TROTTER and FLOTT W WILKINSON (to Minerals Separation North American Corp) U S 1,795,100, March 3 The material is deslimed and then agitated and aerated in the form of a pulp with oil such as pine oil and a compd such as Na oleate adapted to control selective oiling of the phosphate so as to effect froth flotation concn.

Purifying graphite. E RABETRANO Brit 337,738, Nov 9, 1928 Graphite in minerals of low C content is sepd from materials such as silica, silicates or Fe oxides by fractional pptn in a liqnd contg a colloid (such as water with an addn of dextrin, alumina, saponin, gelatin or K silicate), with subsequent treatment of the graphite by electroosmotic action in a vessel in which the settlement of all suspended particles is retarded either mechanically (as by an agitating propeller with upward thrust) or by the presence of a colloid or by both methods App and various details of procedure are described

Gold and silver solutions. A. MOZER. Brit. 338,383, Jan. 7, 1929. Solns. of Au and Ag, suitable for use in refining the metals electrolytically, for treating ores or for electrolytic or other gilding and silvering processes, are prepd by treating the metal with a soln of I in solts such as alkali or alk earth metal iodides, sulfites or thiosulfates forming complex compds with Au and Ag iodides. Chloride or bromides also may be added, and in some instances a cleansing agent such as Vienna lime may be added to form a pasty compn.

Hydrogen for use in catalytic synthesis. LUIGI CASALE. U S 1,793,677, Feb. 24, 1929. CO is produced from CO_2 , O and a carbonaceous fuel contg S, and the CO is subsequently converted by the action of steam in the presence of a catalyst to produce a mixt. of CO_2 and H and the CO_2 is finally removed. In this procedure the conversion of the CO_2 into CO is effected at a temp. (suitably 1000° or higher) high enough to effect the oxidation of any contained S into SO_2 , and the latter is removed from the CO before its further treatment.

Sulfur. WERSCHEN WEISSENFELSER BRAUNKOHLEN A-G. Ger. 514,570, May 8, 1928. S is prepd from gases contg H_2S by leading it over a contact agent of large surface in which excess of SO_2 has been adsorbed. The SO_2 oxidizes the H_2S with liberation of S.

Sulfur from ammonium sulfide solutions. I G FARBERIND A-G. Brit. 337,395, July 29, 1929. Decompn of NH_4 sulfide solns at temps above the m p of S is effected in the presence of adsorbent substances such as silica gel, $\text{Al}(\text{OH})_3$, $\text{Mg}(\text{OH})_2$, $\text{Zn}(\text{OH})_2$, ZnS or other finely pulverized substances or raw cellulose fiber (suitably countercurrent-wise with steam under 1.5 atm. pressure).

Active carbon. I. G. FARBERIND A-G (Ludwig Klebert, inventor). Ger. 517,428, Feb. 22, 1928. Subdivided wood is stirred for some time with a hot soln of an activating agent, e. g., ZnCl_2 or H_3PO_4 , or boiled with the soln, and the mixt. is then dried and heated as usual. The wood may first be swollen with hot water. Cf. C A 25, 1045.

Activating carbon. I. G. FARBERIND A-G (Heinz Thienemann and Julius Drucker, inventors). Ger. 517,316, Mar. 12, 1926. In activating C by means of superheated steam or CO_2 , an activating vessel having a wall permeable to gases is used and the activating agent is caused to diffuse therethrough. Thus, activation may be effected in a permeable chamotte tube, mounted concentrically in an iron tube, into which the activating agent is introduced.

Reactivation of activated carbon after use. C. H. LANDER, F. S. SINNATT, J. G. KRVO and W. E. BAKES. Brit. 337,348, July 23, 1929. After employment for absorption of S compds such as H_2S and CS_2 , reactivation is effected by heating and treatment with a basic substance such as NH_3 , aniline, pyridine or quinoline (which may be used with superheated steam and a small proportion of air).

Carriers for catalysts. IMPERIAL CHEMICAL INDUSTRIES, LTD., F. A. F. CRAWFORD and W. A. P. CHALLINOR. Brit. 337,761, May 3, 1929. Catalysts for SO_2 production and other uses, such as Pt, V and Cr compds, are used on a carrier produced by incorporating "guhr" with a soln of a sol. silicate, pptg gelatinous silicic acid from the soln as by the action of an acid or of $(\text{NH}_4)_2\text{CO}_3$ and drying to obtain a hard porous mass.

Pressed condensation products. SOC. ANON. FOUR L'IND CHIM. A. BALE. Swiss 142,936, Oct. 23, 1928. Filling material, aromatic amines, aldehydes and acid are intimately mixed to form a condensation product which is treated with an acid fixing agent. The product is dried and pressed. Thus, PhNH_2 , CH_3O and HCl are mixed into an aq. suspension of asbestos. NaOH is added and the resulting mass dried, and pressure-molded to the desired shapes. Other examples are given.

Condensation products of fatty acids with polyalkylene-polyamines. I. G. FARBERIND. A-G. Brit. 337,368, July 26, 1929. Fatty acids such as palmitic or oleic acids or their esters such as fatty oils or fats are condensed with diethylenetriamine, triethylenetetramine or a higher polyethylenepolyamine or polyalkylenepolyamine or mixts of such compds. If not sol, the products may be alkylated and they may be used as wetting, emulsifying, foaming and washing agents and are stable in acid, neutral and alk. solns. Several examples are given.

Condensation products of urea and formaldehyde. MARTIN LUTHER, WILHELM PUYGS, ROBERT GRIESBACH and CLAUD HEUCK (to I. G. Farbenind. A-G). U. S. 1,794,084, Feb. 24, 1929. SO_2 is used for hardening viscous completely condensed condensation products of urea and CH_2O .

Polymerized vinyl alcohol-aldehyde condensation products. I. G. FARBERIND. A-G. Brit. 337,800, Aug. 6, 1929. Condensation products of rubber-like, plastic

and elastic or hard and brittle character are made by treating polymerized vinyl alcs. (which may be produced by saponification of polymerized vinyl esters) with aliphatic aldehydes such as formaldehyde or acetaldehyde, in the presence of acid catalysts such as HCl and H_2SO_4 . Several examples with details of procedure are given, and products obtained may be formed into sheets, films, threads, etc.

Objects of albuminous material. INTERNATIONAL KUNSTHOORN INDUSTRIE N. V. (U.S. 514,771, Aug. 3, 1929). Objects of insol albuminous material of a casein character, especially animal casein, are formed by molding the objects from pure casein, subjecting them to a CH_3O bath without allowing them to harden throughout, drying the surface at 100° and molding to final form by cold pressure.

Impregnating compositions, adhesives. I. G. FAABERVIND A. G. (Wilhelm Pungs inventor). Ger. 517,279, Dec. 9, 1926. A mixt. of stearin pitch or like fatty pitch and vegetable or animal oils is heated with S or other sulfurizing agent, e. g., S_2Cl_2 , and before, during or after this treatment a product obtained by extg. tar with a partial solvent, e. g., LiOH , is added to the mixt., heating being continued until the product has the desired consistency. Asphalt, tar, pitch, resins or fillers may be added to the mixt. or to the product. The animal or vegetable oils may be preliminarily oxidized or polymerized. Examples are given.

Fly and caterpillar adhesives. BERNHARD FIBES (Viktor Scholz, inventor). Ger. 514,727, May 23, 1925. Adhesive material for catching flies and caterpillars consists of a 1-15% soln. of cellulose esters or celluloid waste products in castor oil. Odorants coloring matter or filling agents may be added.

Penicils, crayons and chalks. A. FISCHER and P. FISCHER. Brit. 337,633, Nov. 27, 1929. A sulfonated oil such as Turkey red oil is added to the mixt. used, before baking it.

Material for magnet cores, loading coils, etc. GENERAL ELEC. CO., LTD., and G. R. POLOREEN. Brit. 338,321, Oct. 24, 1929. Powd. magnetic material such as an alloy of about 80% Ni and 20% Fe (which may be prepd. as described in Brit. 327,865 (C. A. 24, 5015)) is heated in a soln. formed of chromic anhydride and a halogen acid or halogen salt such as NaCl , preferably washed, and subjected to high pressure. Powd. refractory material such as ball clay, MgO or Al_2O_3 , may be added to reduce the permeability of the core and increase the insulating of the core particles, and the cores after being formed under high pressure may be heat treated to develop their final permeability and hysteresis value (suitably at $500-600^\circ$). If desired, sol. matter may be left on the material until after the pressing and dissolved out after the heat treatment by a prolonged washing.

Wax emulsions. C. H. THOMPSON and W. J. MCGIVERN. Brit. 338,170, July 9, 1929. Emulsions, suitable for use in polishing compns., soaps, for waterproofing fabrics and in the manuf. of lubricating greases, are made by adding soap to melted mineral or other non saponifiable wax such as paraffin or ceresin, and other materials such as albumin, casein, glue or china clay or amino compds. such as triethanolamine also may be added as may also waterproofing agents such as varnish, linseed oil, shellac, rubber, silicates, etc.

Wetting agents. I. G. FAABERVIND A. G. Brit. 337,774, April 27, 1929. In a process generally similar to that described in Brit. 306,116 (C. A. 23, 4932), no excess of high mol. org. acids is used for the condensation process, and the resulting esters, amides or ester amides are converted into salts with acids (suitably either before or after treatment with alkylating or oxyalkylating agents). Numerous details and examples are given.

Water-soluble cleansing agents. JOHANN TENGLE. Swiss 142,691, June 5, 1929. Raw material capable of swelling is heated in an autoclave with caustic alkali, sulfonated oil or resin and hydrocarbons. Thus, starch is heated with NaOH , sulfonated castor oil and CCl_4 . $\text{C}_{11}\text{H}_{23}$ may be added. The product is a good fat and grease solvent.

Detergent for removing tar, etc. J. H. WYLLIE. Brit. 338,167, Aug. 14, 1929. A mixt. comprising benzene, turpentine, oil of citronella and olive oil is used for removing tar from, and cleansing, surfaces such as coachwork, fabrics, leather or metal surfaces.

Composition for removing grease and tar from lacquered surfaces. JOSEF W. NEYER. U. S. 1,795,134, March 3. Spirits of camphor 1, "citronella" 1 and kerosene 12 parts are heated together, mixed with separately melted paraffin 2 parts, and the mixt. is cooled.

Composition for removing paint. N. C. W. PAINT & VARNISH REMOVER CO., LTD., and T. K. BONNAR. Brit. 337,943, Oct. 19, 1929. A paint remover is prepd. by cover-

ing quicklime with soft soap and water, stirring the mixt and then successively adding Na_2CO_3 , NH_4 carbonate and $\text{Na}_2\text{S}_2\text{O}_3$.

Removing paint from painted surfaces. J H GRAVELL Brit 337,461, Aug 2, 1929 See U. S. 1,744,463 (C A 24, 1478)

Cleaning agent for type metal. IRWIN HOFMANN Ger 514,726, Jan 3, 1928 The prepn consists of a reducing mixt of wood charcoal borax and earth alkali sulfide (except MgS) Molten colophony, NH_4Cl , NaOH and Na_2CO_3 may also be added

Smoothing etched printing cylinders. CONZETT & HUBER Swiss 141,344, June 25, 1929 Etched Cu cylinders are coated with an anti-etching prepn which is wiped off so as to remain only in the depressions, and the cylinder is subjected to the action of an etching agent till the depressions disappear

Removing the raised parts of etched copper printing cylinders. CONZETT & HUBER Swiss 142,182, July 4, 1929 The cylinders are given a coating of a non-conductor of electricity which is mechanically wiped off, so that the raised parts are bared and the depressed parts covered The cylinder is next placed as anode in a galvanic bath till the raised parts disappear The cylinder is then removed and mechanically smoothed and polished

Offset printing for copying documents, etc. W L LAWRENCE Brit 337,455, Aug 2, 1929 An original is prepd with an aniline or other copying ink, an impression from this is taken onto a rubber or rubber surfaced member previously treated with glycerol or the like and transfer to blank sheets is effected The glycerol may be mixed with an aq or alc soap soln

Waterproof coatings containing rubber latex. PAUL MEYERSBERG Swiss 142,172, Aug 15, 1928 A coating for materials contains natural or artificial rubber latex in a stable semi pptd thickened state This state may be produced by addn of weak org acid Alkali and saponifiable oils and fats may be present In an example, the mixt. contains latex 65, coned NaOH 1, pptd chalk 20, castor oil 5, PhOH 0.5, colophony 4.5 and rapeseed oil 4% Cf C A 24, 2000

Waterproofing coating suitable for use on gas or oil pipes, etc. LESTER KIRSCHBRAUN (to Flintkote Co.). U. S. 1,794,522, March 3 A surface such as metal pipe to be protected from corrosion is treated with an aq bitumen pitch type dispersion, the coating is allowed to dry and form a film, and over this is applied an aq slurry of hydraulic cementitious material which is permitted to dry and form an outer protecting coating.

Wood effects on metal panels. HERMAN C. MILLER (to United Metal Products Co.). U. S. 1,794,523, March 3 A design is etched on a metal surface such as Cu or brass and the crevices and recesses in the etched surface are filled with a plastic mineral inlay such as a mixt. of whiting, graphite, varnish, lime, white lead, bronze and Al powder, water, Pb_2O_3 , and oil and the inlay is baked, then ground flush with the metal surface of the metal, colored with a "chemical coloring" and the surface of the inlay is grained in imitation of wood grain

Use of rubber and cork together in forming surfacings for power pulleys. CHARLES R. GRIFFITH. U. S. 1,793,927, Feb 24 Various details are described, involving the use of granular cork embedded in a lesser wt. of rubber

Material for gramophone records. A. O. THOMAS Brit 337,437, July 1, 1929. A raw record material comprises gelatinized cellulose acetate with camphor or camphor substitute material, dried at 40–80°, the drying time being about 24 hrs for each 0.1 mm thickness Various details of manuf. are described

Material for sound records. H T. BEANS (to Dunum Products Corp.) Brit. 337,796, April 2, 1929 A core such as cardboard made from rag stock is coated with a resorcinol-formaldehyde condensation product in which a plasticizer such as glycerol is incorporated to avoid brittleness Various details of manuf. are described

Material for playing-cards. HARRY L. POHS U. S. 1,794,866, March 3 Cards are formed with a woven textile center fabric, an opaque phenolic condensation product on both sides of the fabric and in its interstices, and a layer of white phenolic condensation product over both sides of the opaque compn.

Dust bag material for suction cleaners. HOOVER Co Brit. 338,253, Sept. 6, 1929. Material such as porous paper formed from pure manilla hemp fiber is strengthened by application of a resin and plasticizer such as a vinyl synthetic resin dissolved in ethylene dichloride or other solvent with dibutyl phthalate.

Preserving and improving gut for tennis rackets, etc., by treatment with synthetic resins. E S BARRELET. Brit. 338,223, Aug. 20, 1929 Either phenol formaldehyde condensation products or those from glycerol and phthalic anhydride may be used Various details and examples of treatment are given

Fire extinguishers. I. G. FARBENEND A-G (August Ruppert, inventor). Ger. 514,925, Sept. 21, 1928. The formation of COCl_2 in fire extinguishing by the aid of CCl_4 is diminished by adding NH_3 and aryl- or alkyl phosphates to the extinguishing agent. Tricresylphosphate is mentioned in examples. Cf. C. A. 25, 1348.

Apparatus for producing fire-extinguishing foam from water and foam-forming powdered materials. HANS BURMEISTER (to Pyrene Minimax Corp.) U. S. 1,792,810, Feb. 17. Structural features.

Fire-extinguishing compositions. DAVID J. BLOCK (to Du Gas Fire Extinguisher Corp.) U. S. 1,793,420, Feb. 17. See Brit. 319,372 (C. A. 24, 2564).

Fire-extinguishing composition. LODIAS J. DU GAS (to Du Gas Fire Extinguisher Corp.) U. S. 1,792,826, Feb. 17. NaHCO_3 is used with smaller quantities of CaCl_2 , Na borate, NaCl and a filler such as infusorial earth, asbestos dust, silica or ash.

Fire-extinguishing system suitable for use with carbon dioxide. SWINSCOURT MEASURE CO. Brit. 338,197, Aug. 15, 1929. Structural features.

19—GLASS, CLAY PRODUCTS, REFRACTORIES AND ENAMELED METALS

G. E. BARTON, C. H. KERR

The employment of boric oxide in the glass industry. F. H. ZSCHACKER. *Glass-Industrie* 59, 846-9, *Chem. Zentr.* 1930, I, 879— B_2O_3 shortens the duration of fusion and clarification, improves the resistance against hydrolysis, the workability in the blowing process and the behavior with regard to expansion, strength and hardness. The quantity added should not be excessive.

The structure of chilled glass. P. LAZAREV AND S. LIDSNYANSKAYA. *Compt. rend. acad. sci. U. R. S. S. Ser. A*, 1929, 107-8, *Chem. Zentr.* 1930, I, 630—The decrease in d of chilled glass is studied from the surface to the center. When plates of chilled borax glass were ground so as to remove 58%, the sp. gr. of the remaining part of the glass fell from 2.286 to 2.280 and n fell from 1.5170 to 1.5158. When 90% by wt. was removed, n became 1.5151. The exam. in polarized light yielded characteristic interference figures, which give a good illustration of the decrease in hardness. No change in sp. gr. (2.284) was observed when the borax glass was cooled slowly.

Optical properties of didymium glass. E. J. MOORE, H. S. LEIN AND D. E. SHARP. *Glass Ind.* 12, 49-51 (1931)—This glass is composed of approx. 49.5% SiO_2 , 14.0% K_2O , 6.0% Na_2O , 20.5% B_2O_3 , and 10.0% impure com. didymium oxide. It has a high constringence, n of 1.53, and a sharp absorption band, but it does not show anomalous dispersion.

The silvering of glass. DONALD E. SHARP. *Glass Ind.* 11, 273-6 (1930)—Although a wide range of formulas may be used, each one is successful only when followed very closely. An ammoniacal Ag soln. is mixed with some reducing soln. such as sugar, glucose, Rochelle salts, etc. Improper cleaning of the glass and excess NH_3 in the Ag soln. are the most common causes of failure. After cleaning, the glass surface should be treated with dil. SnCl_4 soln. and rinsed with H_2O .

Crystal modes of the technical calcium sodium silicate devitrification and their interpretation. HANS JERSEN MARWEDEL. *Sprechsaal* 62, 715-7, 735-9, 753-6, 773-6, 791-3, *Chem. Zentr.* 1930, I, 878-9, cf. C. A. 23, 2003—The forces active in the crystallization process surpass the "crystallizing ability" by the amt. due to the inhibition on account of the internal friction. This more comprehensive term is called "crystallizing tendency." Many photographs of cristobalite and wollastonite crystals are given. The crystal modes are influenced by external factors, therefore they vary extensively. It is shown by means of crystals of an org. glass that the directed crystn. forces of the skeletons can be altered by external factors. In this case the arrangement of the crystal pattern assumes a medium position. The modes of devitrification make it possible to draw important conclusions with regard to errors in the method, which otherwise can be detected only by difficult and time-consuming investigations. It was found that the cristobalite crystals in the glass are subject to an unexplained torsion force which is perhaps to be attributed to an inhomogeneous melt. The hexagonal star-shaped crystals hitherto described in the literature as tridymite are cristobalite crystals. The surroundings of the crystals are so strongly altered by the formation of crystals that this secondary glass sets another phase. The wollastonite-cristobalite paragenesis thus found was followed and interpreted through all steps. Thus by crystn. the glass

is fractionated into residual glasses, each of which assumes a new position in the equil. diagram G SCHWOCH

Quinquennial review of the mineral production of India for the years 1924 to 1928. Clays. E L G CLEGG *Records Geol Survey India* 64, 371-9(1930)—India produces china clay, fire clays, fuller's earth and clay for pottery, tile, etc. A H. EMEAY

The electrical dewatering of clay suspensions. CARL E CURTIS. *J Am Ceram Soc.* 14, 219-63(1931)—A history and full review are given of much previous work. Dewatering by electrophoresis is generally not more efficient than by using the filter-press, but on very fine-grained clays it is superior Data and bibliographies are given, also a short biography of Count Schwerin C H. KERR

Manufacture of magnesia and magnesite brick. FRIEDRICH REINHART *Tonind.-Ztg* 55, 94-7(1931)—Recent German and foreign patents are reviewed ROBERT FELTON FERGUSON

Effects of autoclave treatments on ceramic bodies and clays. H H. HOLSCHEER. *J. Am Ceram Soc.* 14, 207-18(1931)—Porous ceramic bodies increase in vol and wt. in autoclave treatment with steam pressure of 150 lb per sq in The change is rapid at first but tapers off with increasing time Typical American and English china clays also increase in vol, but ball clays do not Prolonged drying treatments after the autoclave test also were studied Data indicate that the vol changes are not necessarily caused by rehydration of the clay within the body Bodies contg clay, feldspar and flint show a greater vol increase than do the pure clay constituents The compn of the body is more important than the absorption in detg the reaction to autoclave treatment C H KERR

The ceramic industry of Ontario. ROBERT J MONTGOMERY *Ann Rept Ont Dept Mines* 39, Pt. IV, 196 pp (1930)—The compn and properties of clay, and the manu of clay products are discussed, and a geological description of Ontario clays, and brief descriptions of all ceramic plants in Ontario are given The latter produce chiefly heavy clay products, but lesser quantities of refractories, pottery, floor tile, elec. porcelain, glass and enameled Fe ALDEN H. EMEAY

Tests on ceramic materials O KALLAUNER *1st Communications New Intern Assoc. for Testing of Materials B*, 263-77(1930)—A standard procedure is proposed comprising specifications for the prepn of samples, mech analysis, tests for fireproof qualities and tests on the prepd. material with regard to its applicability for prepn. as a plastic mass and on the test pieces made from such a mass The latter tests comprise the detn of the amt. of water necessary to work up the sample of earth into the plastic mass, the forming of the test pieces and the detn. of their properties after drying and burning The following properties are detd: behavior while working up to the plastic mass, shrinkage on drying, tendency to form blisters, bending strength, loss of wt. after burning, color, ring, shrinkage on burning, capacity to absorb water and sp. gr. E J C.

Aging and non-aging ceramic bodies. FELIX SINGER. *Keram. Rund* 38, 167-71, 183-7, 216-22(1930)—The failure of porcelain after repeated heating and cooling is due to differential thermal expansion either of (1) the constituents of the body itself, (2) the old constituents and the new, produced by gradual crystn of the glass in the body, or (3) the body and the glaze Disruption due to dissimilar expansibilities of glaze and body may be avoided by using glazes with high elastic limits and high strength. The most satisfactory porcelain for general purposes is one composed of a homogeneous, unstrained glass contg many very fine mullite crystals Steatite porcelain is superior to ordinary porcelain in resistance to aging and in all other important properties except thermal expansion, because it consists of a single constituent. H INSLEY

The influence of time on the maturing temperature of whiteware bodies. I. F. H. NORTON AND F. B. HODGSON *J Am Ceram Soc* 14, 177-91(1931).—Six typical com. bodies showed similar shrinkage and porosity curves for different firing times, but the longer the time the lower the curve was shifted on the temp scale In all cases the usual log law relating temp with rate of reaction applied The const., however, was not the same for all bodies II. CHARLES F NORTON, JR *Ibid* 192-206—With samples that showed the same degree of maturity by phys. tests, the long-fired, low-temp samples showed more quartz soln as well as more mullite development than samples fired more quickly at higher temps. C. H. KERR

The origin of pin holes in cast ware. GERHARD BUDEWIG *Keram Rund* 38, 215-6(1930)—An important cause of "pin holes" is the high viscosity of slips which does not permit the release of air bubbles introduced during blunging A simple pipet viscometer is adequate for plant measurement of viscosity, a relative value for viscosity being obtained by the ratio of discharge time for the slip to the discharge time for pure

Thermostatic control of temperatures in glass kilns, etc. **PILKINGTON BROS., LTD.**, and **P M Hogg** Brit 337,803, Aug 2, 1929 Variations in cond of the material being heated are utilized for temp regulation by an elec system, various details of which are described

Apparatus for "frosting" glass articles such as bulbs. **JAMES BAILEY** (to Corning Glass Works) U S 1,793,893, Feb 24 Treating fluid is supplied under atm pressure to a partially exhausted chamber, over openings in which the articles to be treated are supported Various structural details are described

Glass for use as a light filter. **F WEIDERT** (to Deutsche Gasglühlicht-Auer Ges.) Brit 338,334, Dec 15, 1928 Glass colored with neodymium oxide is used for light filters for accentuating color contrasts Ce oxide may be added to absorb ultra violet rays, or Cr oxide or Pr oxide may be used to make green more pronounced in appearance The glass may be used for spectacle lenses or otherwise for filtering or special illuminating effects

Glass permeable to ultra-violet rays **FRITZ WOLLNER, RICHARD WOLLNER** and **FERDINAND NIKOLAI** Austrian 120,674, Oct. 15, 1929, Fr 691,805, Mar 12, 1930 Glass is rendered permeable to ultra violet rays by reducing the Fe_2O_3 present in the mixt or in the melt to the metallic state This may be effected by addn of Al, Si, Mn, Zn or Ca

Apparatus and procedure for uniting components of "safety glass" assemblies by heat and pressures. **DUPONT VISCOLOID CO** Brit 337,547, Oct 27, 1928

Drying clay ware, etc., in tunnel kilns. **FRANK M HARTFORD** (to Harpor Ceramic Service Co.) U S 1,793,086, Feb 17 Various details of app., humidity and temp regulation are described

Composite brick. **EDWIN M WYATT** (to American Face Brick Research Corp.) U S 1,794,572, March 3 A building brick is formed with a burned facing section of argillaceous material and a concrete backing section formed of an aggregate of prepd burned argillaceous particles and a cementitious binder, this binder serving also to unite the 2 sections

Decorating bricks, flower pots, etc. **A F BERAY** Brit 338,147, Aug 8, 1929. Articles of clay or composite material such as concrete are decorated by facing them with irregular fragments of stone, granite chippings or the like, suitably after initial application of a thin moist layer of portland cement, or by pressing into the surface of a green molded brick

Apparatus for screening potters' slip and other viscous materials. **MASCHINEN-FABRIK VORM G DORST A-G** Brit 337,371, Nov 20, 1928 Structural features

Tunnel drier for pottery. **MÖLLER & PFEIFER** Ger 517,167, Jan 20, 1927.

Magnetic separator for treating porcelain sludges, etc. **MAGNET-WERK GES EISENACH SPEZIALFABRIK FÜR ELEKTROMAGNETAPPARATE** Brit 337,759, April 5, 1929 Structural features

Refractory materials. **METALLGES A-G** Brit 337,605, Nov 7, 1929 Mg silicate materials such as olivine, serpentine or talc contg Fe are heated, without fusion, with substances rich in Mg, such as MgO or magnesite, in an oxidizing atm for the conversion of the Fe present into Mg ferrite and of the free SiO_2 into Mg orthosilicate Materials contg Ni, Mn, Cr and Al may be similarly treated, and numerous details and examples are given

Vitreous material suitable for casting or hot pressing in molds. **PERCY B CROSSLEY** (to Mycalex (Parent) Co., Ltd.) U S 1,795,200, March 3 A mixt formed of colloidal mica, colloidal asbestos, refractory ceramic material such as porcelain or glass infusible below 900°, and metallic flux material such as PbO and H_2BO_3 , fusible below 700° is molded and vitrified Cf C A 25, 573

Tile-glazing apparatus. **GEORGE F WILDE** and **BENNET K. ESKESSEN** U. S. 1,792,788, Feb 17. Structural features

Muffle furnace for enameling sheet-iron ware. **MANFRED WEISS STAHL- UND METALLWERKE A-G** Swiss 141,563, May 21, 1929

20—CEMENT AND OTHER BUILDING MATERIALS

J. C. WITT

The testing of cements by means of earth-moist and plastic mortars. **J P LOEVEN-THAL**. 1st Communications *New Intern Assoc for Testing of Materials B*, 204-9 (1930) — In earlier comparative investigations carried out by the Danish state testing house the Swiss prism method and the common hammer method were used, in continuation,

a series of tests was made dealing with mortars of portland cement, aluminous cement and a Danish pozzuolana cement (Moler cement). The test pieces were made in compact molds and were of consistencies ranging from earth moist to fluid. They were tested after 7 and 28 days. It was shown that at both times the strength curves of the portland and Moler cements were quite similar, although there was no proportionality. In contrast with this, mortars of aluminous cement of low water content showed an increase in strength between 7 and 28 days, while mortars of high water content showed a fall in strength. The various methods of testing are discussed. **I. J. C.**

Some properties of high-alumina cements from six countries. **P. H. BATES. 1st Communications New Intern Assoc for Testing of Materials B, 210-8(1930)**—High-alumina cements (ciment fondu) made in 6 different countries were exam'd for their properties by chem., petrographic and phys. methods. The phys. characteristics were det'd by strengths in the form of 1 3 standard Ottawa sand tensile briquets and 2-inch compression cubes. Concrete was also tested in the form of 6 X 12 in. cylinders made from a 1 2 4 mix of gravel. The chem. analyses showed the cements were rather widely different from one another. The petrographic exam'n. showed that the alumina was largely present either as $\text{CaO} \cdot \text{Al}_2\text{O}_3$ or as the unstable form of $5 \text{CaO} \cdot 3\text{Al}_2\text{O}_3$. The wide variation in comp'n. and the difference in constitution did not result in as great a difference in phys. properties as would be expected. The effect of the heat generated during the setting of the cement and the effect of the amt. of mixing water used were also studied. Increased temps. during setting resulted in decreased strength. Increased amts. of mixing water resulted also in marked decrease in strength. **L. J. C.**

Hydraulic additions and trass-portland cement. **S. I. DUCHENOV. 1st Communications New Intern Assoc for Testing of Materials B, 221-37(1930)**—Volcanic tuff from the Crimea and diatomaceous earth from tested sources, when ground to the fineness of portland cement and added to portland cement clinker, give pozzuolana portland cement which, when used in hydraulic structures, is more resistant to sea water and acid water than is portland cement mortar. Its mech. strength after 3 months' hardening in water is equal to that of ordinary portland cement mortar without hydraulic add'n, the same quantity of sand being used. The best methods for rapidly detg. the degree of activity of the hydraulic add'n. are to est. the amt. of lime it absorbs and to compare the strength of mortars, with and without such add'n, under the conditions afforded by accelerated hardening of the test pieces at 80°. The best method for detg. the optimum add'n. of hydraulic material is the above-mentioned accelerated or thermal method. Pozzuolana portland cement produces a more voluminous paste than ordinary portland for the same quantity of sand. **E. J. C.**

Fine grinding of cement. **ERNST RUSSEL. Zement 19, 1079-80(1930)**—The water permeability of concrete is more greatly affected by changes in the fineness of cement than is the compressive strength. The grading of the finer fractions of the cement also has a definite bearing on its ability to make concrete impervious. **H. F. K.**

Mill temperatures and the setting time of cements. **K. KOYANAGI. Zement 19, 988-9(1930), cf. C. A. 24, 5452**—Cements high in SO_3 were found to set rapidly under high mill temps. while those of low SO_3 content were not affected. **H. F. KRIEGER.**

Portland cement materials in Arkansas. **W. R. SPENCER. Univ. Arkansas, Eng. Expt. Sta. Bull. No. 6, 27 pp(1929).** **E. H.**

Investigation of portland cement and its constituents by means of vapor-pressure measurements. **F. KRAUSE AND G. JÖNSS. Zement 19, 1054-5(1930)**—The use of a micro-tensi-eudiometer is proposed to follow the formation of cryst. or colloidal phases in the hydration of cements. **H. F. KRIEGER.**

Free lime in portland cement and soundness. **G. HAEGERMANN. Zement 19, 982-4(1930)**—Under a range of burning temps. of 1225-1475° the free lime in 12 cements was found to vary irregularly with temp. from 0.24 to 1.85%. In general, cements were sound in cold H_2O when they contained up to 4% free lime, while 1-2% was the upper limit when subjected to the boiling test. **H. F. KRIEGER.**

Free lime, soundness and strength (of cement). **A. GUTTMANN. Zement 19, 1078(1930)**—G's investigations confirm those of Haegermann (cf. preceding abstr.) in regard to the content of free lime and soundness in cements. Further G. found a definite increase in tensile strength with a decrease in the free CaO present in the cement. **H. F. KRIEGER.**

Action of gypsum, swelling and hydration. **P. TIFFMANN. Zement 19, 1030-5, 1056-60, 1080-6, 1106-12(1930)**—The properties of gypsum, raw and in its various phases of dehydration, are discussed in relation to their effects on cement. Numerous photomicrographs are given showing the cryst. products of the hydration and hardening stages of cements and other materials. **H. F. KRIEGER.**

Quicker tests of cement and concrete. H. KRÜGER. *1st Communications New Intern Assoc for Testing of Materials B*, 114-8(1930)—Preliminary tests made to investigate whether it might be possible to decide quickly (in 1 or 2 days instead of 28 days) whether a cement is fit for delivery showed that there probably is a fairly good concordance between the normal crushing strength of cubes after 28 days' combined air and water-curing and after a few hours' steam hardening. The method may be of some use also for quicker testing of concrete, but has not yet been satisfactorily developed. F. J. C.

Compression, tension and bending tests, cement and concrete. A. T. GOLDBECK. *1st Communications New Intern Assoc for Testing of Materials B*, 61-6(1930) F. J. C.

Cement, concrete and reinforced concrete—laboratory tests and control and practice on the building site. A. HASCH. *1st Communications New Intern Assoc for Testing of Materials B* 91-9(1930) (in German), cf *C A* 24, 3874—An attempt is made to consider briefly, from the present day standpoint, the relative merits of ordinary and high-strength mortars and concrete, and tests to which these materials should be subjected in the lab. and on the building site. The essential point is organization of control tests on the site, combined with chem. lab. tests, because samples taken before building is started are of no value. The theory of probability is applied to results of static tests. E. M. SYMMES

Resistance of concrete pipes to corrosion by water. J. O. ROOS. *1st Communications New Intern Assoc for Testing of Materials B*, 111-8(1930)—In Sweden decomposition of concrete pipes is due mostly to the external action of H_2O percolating through the pipe wall and extg. lime from the cement. Pure H_2O has a high disintegrating effect, a water contg. 0.30-80 mg. per l. of CO_2 increases and lime hardness decreases the corrosion. To be durable, concrete pipes must be impermeable to water. E. M. S.

Slag concrete. I. BENTKOVSKII. *1st Communications New Intern Assoc for Testing of Materials B*, 234-43(1930) (in German)—Tests were made by the Institute of Metals, Leningrad, to use a particular slag for concrete. E. M. SYMMES

Resistance of concrete to chemical attack. K. MANDL. *1st Communications New Intern Assoc for Testing of Materials B*, 149-56(1930)—In Austria the deterioration of concrete due to injurious influences is relatively slight. The most prominent instance is the destruction of a pillar of a new dwelling house in Vienna, which was definitely found to be due to the action of gypsum. It is especially noteworthy that this damage was not due to sulfate-contg. ground water, but to surface water that had percolated through the slag-contg. bottom layer, from which it dissolved the sulfate, and then was forced to pass through the concrete pillar, because the slag deposit was laid on material impermeable to water. Several cases of deterioration of the sewer system of Vienna by sulfate-contg. water are also reported. Results of an investigation into the action of nitrates on concrete have been published in part, but it is not yet possible to draw conclusions from this work. E. J. C.

The compactness of concrete and its resistance to chemical action. A. POULSEN. *1st Communications New Intern Assoc for Testing of Materials B*, 157-61(1930)—Even the best concrete is destroyed by the reaction of its free lime with the $MgSO_4$ contained in sea water. Rich mixes soon show cracks of their own accord. The application of a protective coating to large masses of concrete employed in marine work is, of course, impossible. To neutralize the excess lime it is necessary to add to the cement pozzuolana contg. in sufficient quantity the necessary hydraulic constituents (e. g., SiO_2 sol in alkalis). The cement with hydraulic admixts. thus obtained should be submitted to lab. tests. In Denmark, the Danish diatomaceous earth, Mo Ler, which contains about 60% of sol. SiO_2 , has been used with considerable success for the past 20 years. F. J. C.

Evaluation of shrinkage data on light-weight concrete. A. HUMMEL. *Zement* 19, 1062-6(1930)—While light-weight concretes have greater vol. changes than do concretes contg. the usual mineral aggregates over the same humidity ranges, these vol. fluctuations need not imply greater danger of crack formation. H. F. KRIEGER

Method of disintegrating hardened concrete for the determination of the original ingredients. R. LOMAY. *1st Communications New Intern Assoc for Testing of Materials B*, 167-79(1930)—It is possible to disintegrate the concrete satisfactorily without smashing up the aggregate by heating for a time to 800-900° and cooling by suddenly quenching in water. The aggregate is sepd. from the cement and can be sieved, it is almost completely recovered, but has generally become a little finer as a result of the ignition and embrittlement on quenching. This increase in fineness is revealed by the sieving. The proportion of cement can be estd. by various methods. Expt. has proved, however, that the most rapid method and the one that gives the most reliable results

concrete heating to 800-900° (which removes combined water and CO_2) and then cooling in a space free from water vapor and CO_2 . The wt. obtained is that of the original cement less its loss on ignition. The method is of sufficient accuracy for technical work. E. J. C.

Deterioration of concrete in hydraulic structures. A. FRUWALL. *1st Communications New Intern. Assoc. for Testing of Materials B*, 162-6(1930). *Id. C. A.* 23, 5558—Inspections of hydraulic structures in Sweden 10 years ago showed some defects in the concrete and continued observations proved that the concrete in some cases was subject to an increasing deterioration, apparently due to the quality of the concrete materials, to the methods of making concrete or to the conditions to which the concrete was exposed. The aim of these investigations was to ascertain the causes of such increasing deterioration. It had been observed that deterioration occurred only in structures exposed to water on one side and was caused by the soly of the lime and the decomposition of other chem. combinations in the hydrated cement. Particular attention was therefore directed to the soly of the cement and the aggressiveness of the water. Practical expts. were made at the same time with concrete slabs exposed to low water pressure during several years. It was proved that Swedish natural waters need not be considered aggressive in this respect. The importance of using sand relatively free from org. substances was impressively established. Leaner mixes than 1:4:5 should not be used for concrete exposed to water pressure from one side. Replacing cement by an admixt. of 10% lime or 5% CaCl_2 improved the watertightness, while 25% trass or slate powder gave practically no improvement. E. J. C.

Natural and artificial stone—mineralogical and petrographic properties. Methods of testing. L. FINCKH. *1st Communications New Intern. Assoc. for Testing of Materials B* 21-4(1930). E. J. C.

Coöperation of the mineralogist and petrographer in judging natural and artificial building stone and road stone. P. NIGGLI. *1st Communications New Intern. Assoc. for Testing of Materials B*, 1-9(1930).—The mineralogist should conduct the sampling of the material so as to provide a typical sample for investigation, and thoroughly describe it from the mineralogical standpoint as to its chem. nature, durability, structure and texture. After the tech. tests, he should evaluate the behavior of the material mineralogically and relate it with the properties found. The mineralogical petrographic methods can be conducted with min. quantities of material and are therefore particularly suited for control expts. It is thus important to elucidate experimentally the relation between tech. behavior and the nature of the material, so that tech. tests can be more and more replaced by mineralogical examn. I. J. C.

The application of mineralogical and petrographic knowledge in the testing of non-metallic inorganic materials. R. GRENGG. *1st Communications New Intern. Assoc. for Testing of Materials B*, 13-20(1930). E. J. C.

Structure of weather resisting rocks (VAN DER VEEV) 8. Siliceous CaO compounds [for use in cold cement industry] (Ger. pat. 514,742) 18. Decorating bricks, flower pots, etc. (Brit. pat. 338,147) 19. Rubber compositions [for expansion joints, wall and roof coverings, etc.] (Brit. pat. 338,247) 30. Producing carbide-forming metals [with cement as a by product] (U. S. pat. 1,794,401) 4. Forming sheet insulating material such as wallboard (U. S. pat. 1,794,433) 13.

Concrete Year Book, 1931. A Handbook, Directory and Catalogue of Concrete, 8th year. Edited by O. FABER and H. L. CHILDS. London: Concrete Publications, Ltd. 3s. 6d., net.

HERTEL, PAUL. Einfluss der Verwendung von Edelschlag auf die Güte und die Kosten von Beton. Charlottenburg: Zementverlag. 64 pp. M. 2.

Cement. PERRIN & TUSCHER. SWISS 141,523, April 25, 1929. The cement contains essentially blast furnace slag, lime and bauxite. Gypsum may also be added.

Portland cement. ANDREW LUNDSTROM and GUY O. GARDNER (to Ash Grove Lime & Portland Cement Co.) U. S. 1,794,526, March 3. A quick-hardening cement is made by adding, to a clinker formed from argillaceous and calcareous materials, gypsum and a mixt. of CaCl_2 and NaCl sufficient to act as accelerators for hardening the final product, and grinding the aggregate thus formed with a lubricant such as Al stearate so that at least 75% passes a 200-mesh screen.

Chlorine-treated cement. HANS M. OLSON. U. S. 1,792,755, Feb. 17. Cl_2 gas is directly mixed with hydraulic cement during its grinding in order to facilitate attainment of rapid set and early strength.

Apparatus for testing cement, etc. W. D. WILLIAMS. Brit 337,788, Aug. 2, 1929. Various details of app. are described suitable for use in boiling substances for testing as in the Le Chatelier test for cement, boiling cement pats, testing porosity or permeability of cement disks, etc.

Cement-coloring and hardening composition. MAXIMILIAN TOCH (to Standard Varnish Works) U S 1,793,143, Feb 17. A material for coloring and hardening portland cement comprises a dry mixt including a coloring material resistant to portland cement such as Fe oxide, Cr oxide, MnO_2 or lampblack together with $CaCl_2$ and a colloiddally pptil material such as Al silicate, Bi_2SO_4 or Co aluminate.

Feeding cement material to kilns. POVL T. LINDHARD (to F. L. Smith & Co.) U S 1,793,499 Feb 24. Slurry is delivered to a kiln as an atomized spray conc, dust is sep'd from the kiln gases and this dust is returned to the kiln at a point where the conc of atomized slurry substantially fills the diam of the kiln. App is described.

Rotary kiln (with slack chain groups in its interior) suitable for burning cement-forming slurry. JOHAN S. LASTING (to F. L. Smith & Co.) U S 1,793,471, Feb 24.

Rotary kiln suitable for cement manufacture, etc. HAAVARD KRONSTAD (to Bessemer Cement Corp.) U S 1,793,408, Feb 17.

Concretes and mortars. J. W. BATTERSBY. Brit 338,242, Aug 27, 1929. A waterproof product is prep'd by adding a powder of colloiddally readily sol water-contg alkali silicate 2-5% to portland cement or lime concrete or mortar prior to mixing with water (preferably to the cement clinker before grinding it). The initial set period may be controlled by varying the amt of water in the silicate (which may be 10-30%). For some purposes such as sealing porous places in floors a much larger proportion of silicate powder may be added.

Apparatus for molding concrete blocks. FAYRE & CIE. ZELENYHETONFABRIK WALLISELEN. Swiss 142,688, May 15, 1929.

Bituminous paving materials. EDWIN C. WALLACE. U S 1,793,845, Feb 17. Bituminous paving compns workable at atm temp are prep'd by coating heated particles of mineral matter with heated bituminous cement in sufficient quantity to cause agglomeration when the mass becomes chilled and, before chilling, adding to the mass cold, non-coherent, non hygroscopic material comprising mineral particles coated with bituminous cement, proportioned to render the entire mass granular and friable but capable of consolidation by the application of pressure alone. Cf C A 25, 793.

Bituminous compositions for roads, etc. F. A. HILL. Brit 338,206, Aug 16, 1929. A cold-lay compn suitable for surfacing roads, canal beds, etc. and which is stable when stored in hermetic containers) is made by mixing at a temp of 100° or higher, asphalt, bitumen, tar, pitch or asphaltic petroleum residues with a fuel oil or petroleum product from which all the lighter portions including the illuminating oils have been distd. The material may be used in the proportions of 35-55 and 65-45%, resp.

Bituminous coating compositions. I. G. FARDENING A-G. Brit 337,521, Sept. 4, 1929. Products for covering roads, lining vessels, etc., are made by incorporating bituminous materials such as natural or artificial bitumens, mineral or brown coal tars or their distn products with synthetic polymerization products of diolefins such as those of butadienes, which incorporation may be effected in the presence of a diluent such as benzene or cyclohexanol, and if desired with emulsification with water effected by a suitable emulsifying agent. Various details and examples are given.

Bituminous pavements. JOHN RADCLIFFE (to Colprovia, Ltd.) U S reissue 17,985 March 3. Reissue of original pat. No 1,655,240 (C A 22, 1026).

Road-surfacing materials. I. G. FARDENING A-G. Brit 337,928, Oct 12, 1928. Shippery deposits on roads are removed by use of org solvents or their aq solns or emulsions to dissolve the oils and lats or swell the rubber in the deposits, followed by brushing and washing with water. Various examples of solvents and emulsions are given.

Rubber-surfaced paving blocks. EDWIN C. WALLACE. U S 1,794,220, Feb. 24. Structural features.

Glass paving blocks. PIERRE NOEL. U S 1,795,229, March 3. Structural features.

Material suitable for expansion joints. ALBERT C. FISCHER (to Philip Carey Mfg Co.). U S 1,793,439, Feb 17. Sheets are formed with a waterproof ductile binder such as blown asphalt having mixed with it flat and relatively thin strips of fibrous roofing scrap. Cf C. A. 24, 2574, 4608.

Artificial stone. JERRE HAGGARD. U S 1,793,172, Feb 17. A vesicular aggregate such as cinders or basalt is used with granulated gypsum and hydmted lime.

Artificial marble residues. **ALEXANDER STACE** Ger 514,710, Nov. 19, 1927. The residues from the manu. of artificial marble contg MgO cement are ground and all sol constituents removed by lixivation with water. The residue is then used as flux for portland cement mortar.

Porous building material. **FRIK C RAYPA** U S 1,704,272, Feb 24. A tenacious foam which may be formed from resin soap, gelatin and water is mixed with mortar and the material is formed and cored under high steam pressure. Cf C. A. 24, 4134.

Waterproof sheet material for building construction. **F. RUDIN** and **H. KOLLBRUNNER** Brit 348,403 Dec 13 1928. Corrugated cardboard layers are superposed with the corrugations of adjacent layers at right angles to each other, fastened together by threads or wires or the like without adhesive, dried at 60-70°, impregnated with coal tar pitch or the like and pressed.

Fireproofing binder for use with fibrous material. **CHARLES H. BROWN**, U S. 1,703,357 Feb 17. A compn suitable for use with fiber in making wall board, etc., consists solvly of NH_4 phosphate, H_2BO_3 , dextrin and water, with a greater proportion of NH_4 phosphate than of H_2BO_3 and a greater proportion of dextrin than of NH_4 phosphate.

Roofing material. **RALPH F. DRAKE** (to Anaconda Sales Co.) U S 1,704,449, March 3. Crushed and graded elec cond material such as crushed coke substantially free from dust is applied as a layer to the surface of material such as asphalted roofing felt so as to form an irregular surface, and a cond powder such as graphite is applied over this surface so as to fill interstices between the granules, and a metal such as Cu is applied as an electrodeposited layer over the cond material.

Plaster. **V. LEFFNER** Brit 347,926, Oct 10, 1929. Anhydrite is mixed with plaster of Paris or "any standard plaster" such as Keenes cement, and the anhydrite may be "accelerated" as described in Brit 236,693 (C. A. 20, 1310) and Brit 317,672 (C. A. 24, 2267), and by varying the relative proportions of the ingredients, the setting time may be varied upward from an almost instantaneous set as desired. Various details and examples are given.

Plaster. **FOSTER D. SNELL**, U S 1,792,661, Feb 17. A plaster of "retarded suction" comprises a calcareous binder such as slaked lime, a filler such as sand, together with water, and a "minimal" quantity of dispersed lime soap present to the extent of only about 0.16 lb per cu yd.

Plasterboard. **HAROLD L. LEVY** (to Plimkote Co.) U S. 1,703,810, Feb 24. See Brit 310,213 (C. A. 24, 2573).

Kila suitable for drying lumber. **ALBERT ARMSTRONG** U S 1,703,802, Feb 24. **Preserving wood.** **I. G. LARSEN** and **A. G. (Wilhelm Bonrath, Wilhelm Schepers and Karl Taube inventors)** Ger 515,207, Feb 10, 1928. The wood is treated with compds in which Hg is linked by at least one of its valencies to an unsubstituted or neutrally substituted hydrocarbon radical. Phenylmercuric acetate and ethylmercuric bromide are suitable. The compds may be used in soln in an org solvent, or in aq emulsion or soln and their soly in water may be increased by addn of alk compds. Examples are given. Cf C. A. 25, 1056.

Treating wood. **ADOLF KISTNER, JR.** Ger 514,990, Nov 24, 1929. Wood surfaces are protected by an elastic waterproof layer of bitumen admixed with fine wood meal.

21—FUELS, GAS, TAR AND COKE

A. C. FIELDNER AND ALDEN H. EMERY

Pulverized fuel. **ANON** *Fuel Economist, Pulverized Fuel Suppl* 1-72, Oct., 1930. An illustrated review of plant and equipment, some notable British installations and use in Lancashire boilers in the metallurgical industries, and in marine practice.

The use of sugar for motor fuel. **K. ŠANDRA** and **J. ZEMAN** *Z. Zuckerind. tschechoslov. Rep* 55, 172-4 (1930).—See C. A. 25, 576. **LESLIE B. BRAGG**

Tests on agglomeration of combustibles. **M. G. LEVI**, **C. PADOVANI** and **F. CARDIA** *Atti III congresso naz. chim. pura applicata* 1930, 753-66.—Briquetting tests were made on lignite, anthracite, semicoke, peat and mixts of these with sulfite tar, mineral oil, linseed oil and mixts of these. Resistance to fracture and atmospheric effects were detd. **J. F. LEBRE**

Determination of moisture in combustibles. **C. PADOVANI** and **C. SINIRAMÉ**

Atti III congresso nau chim pura applicata 1930, 778-81—Combustibles may oxidize during the detn of moisture (*Compt rend 7 congresso chim ind Oct., 1927*). All coals, heated to 60-70° in a current of N_2 free from O_2 , evolve small quantities of CO_2 . A method was developed to det simultaneously the loss of wt of the coal by heating at 100° in a current of N_2 and increase in wt of a $CaCl_2$ tube. Loss of wt of the coal is greater than increase in wt of the $CaCl_2$ tube. E. M. SYMMES

Incomplete combustion: its importance, recognition and prevention. E. W. B. DUNNING *Gas J* 193, 271 6, *Gas World* 94, 127-30(1931)—A review is given of the fundamental principles of combustion together with the physiol effects of the products of complete and incomplete combustion. The importance of CO in incomplete combustion is emphasized. It is the only gas that can be present in sufficient concn to be dangerous in the flue products from gas-consuming appliances. The combustion characteristics of various types of gas appliances are reviewed and the amt of CO to be expected is shown to be below any figure at which adverse physiol effects can occur. The products of combustion of town gas can be and should be harmless. When this is not the case the cause can be traced to the various factors in the design of the gas appliances. The degree of incompleteness of combustion is detd by the analysis of the flue gases for CO. The estn of the products of incomplete combustion offers some difficulty because of the small concn in which they are present and special methods must be used. The $CuCl$ method is not suitable for detg CO. The preferred method, which is very sensitive, is the oxidation of CO with I_2O_5 according to the following equation $I_2O_5 + 5 CO = I_2 + 5 CO_2$. A train of app for this method is described and illustrated. An accuracy of 2-3 parts of CO per million of flue gases can be obtained. F. H. BERGEM

The problem of lignite in Italy. ALDO RRGÉ *Chim ind agr biol* 7, 17-27 (1931)—Lignite was found in small quantities in every Italian Province, but mostly in Tuscany. The Valdarno lignite contains moisture 51.8, volatile matter 23.35, ash 6.60, fixed C 15.25%, and has a heating value of 7663 cal (on dry basis). The gas obtained by distn has a heating value of 5592-6192 cal, the coke is of had quality, but a very good coke is obtained by mixing the lignite with 25-30% of ordinary coal. The Piedmont lignite has: moisture 10-13, volatile matter 30-36, fixed C 34-45, ash 12-13%. The xylid lignite from Spoleto has: moisture 35-45, volatile matter 30-45, ash 8-18, fixed C 27-38%, heating value 3-4000 cal and the gas obtained by distn contains: CO_2 12-14, O_2 0.2-0.7, H_2 25-29, CO 11-14, C_2H_4 2-5, N 34-40% and has a heating value of 1050-1350 cal. G. A. BRAVO

The torbanites of South Africa. F. H. CUNNINGHAM CRAIG *J. Inst Petroleum Tech* 16, 620-5(1930)—See *C. A.* 24, 2860. EMMA E. CRANDAL

Some modern ideas on coal. F. V. TROFSWELL *Colliery Guardian* 140, 1651-4, 1744-6, 1845-6(1930)—A detailed summation is given of the present state of our knowledge of the structure and compn of coal. E. M. SYMMES

The classification of Roumanian coals. I. ILIUT *Bul chim soc romand stinte* 31, No 4/6, 3-11(1931)—The basis for the classification of coals is clarified. The geologic age is not a sufficient criterion for the purpose. Physico-chem and petrographic characteristics of the coal must also be considered. The decisive factors in detg the chem age of coal are (1) elementary compn, (2) behavior on distn and (3) resistance to action of certain solvents. The later coals contain less C and more O, have a lower calorific value, and give off more gas (largely CO_2) and at a lower temp than the earlier ones. In consideration of these ideas Lupenl coal, usually considered a superior lignite, should be classified as bituminous. CHANNING WILSON

The role of the components of Roumanian fossil coals in the composition of the gas distilled from them up to 500°. I. ILIUT *Bul chim soc romand stinte* 31, No 4/6, 13-9(1931)—Four Roumanian coals and lignites were extrd with solvents, and the residues from each extrn submitted to distn. Bitumen was removed from the original coal by extrn with a benzene alc mixt, lignin by extrn with H_2SO_4 in an autoclave, and humic acids with NH_3 soln at 90°. No cellulose was found in the coals examd. The residues richer in humic acids and lignin gave more gas on distn. Unsatd hydrocarbons were less in residues after extrn of bitumen. CO and CO_2 were found in larger quantities in gas obtained from residues rich in lignin, and especially humic acids. The gas from the carbonaceous residues after all three extrns contained a larger percentage of C_2H_4 . CHANNING WILSON

The coal industry of South Africa. T. COULTER *J. Chem Met Mining Soc S. Africa* 30, 348(1930); cf. *C. A.* 24, 4371, 5133, 25, 576—Discussion. A. H. F.

Washability studies of the Black Creek bed at the Bradford Mine, Dixiana, Ala. B. M. BIRD, B. W. GANDRUD AND C. B. BARMORE. *Bur. Mines, Rept. of Investigations*

3083, 12 pp (1931) —The prepn of a washed coal contg less than 2% ash should be practicable with a comparatively simple washing system. The coal contains a large proportion of fine sizes due to friability. There are only small quantities of flaky impurities. The coarse and fine sizes should be separately washed. A. H. B.

A contribution to the coal-dust problem. I. Mathematical treatment of the free fall of dust particles. II. Deposition of dust particles from a horizontal gas stream. III. Centrifugal separation of dust from gases. Fa. PROCKAT *Glasters Ann* 106, 73 9, 93-7 151-4, 107, 39-45, 47-51 (1930) —A formula for free fall of dust particles is derived, which is superior to that of Stokes, and agrees with the data within 10%. Mathematical formulas are derived for horizontal settling chambers and centrifugal separators which are in agreement with data from exptl equipment. L. W. T. C.

Determination of sulfur in coal, gas and purifying material. J. H. STINKAMP *Gas Age Record* 67, 53-4 (1931) —The Parr KClO_4 method for coal is briefly described. A method for gas, of burning the gas, oxidizing the SO_2 in 3% soln of H_2O_2 and titrating with 0.1 N KOH is described. For S in purifying material, the material is extd with CS_2 , tarry matter being removed by passage of the CS_2 through a tube filled with activated coal. The CS_2 is evapd in a crystg dish which is placed in a drying oven at 50° for one hr, leaving pure S, which is weighed. LESLIE B. DRAGO

The assay of coal for gas manufacture. J. G. KING *Gas Eng* 48, 75-7 (1931); cf. C. A. 24, 944 —A modification of the Gray-King Method (C. A. 23, 5026) for the assay of coal is described, making it suitable for application to coals used in gas manuf. The volatile products of carbonization are passed through a cracking zone in which the temp is controllable. The probable yields of products are calcd from the assay yields by suitable factors obtained by comparing the yields obtained from a certain no. of typical coals. The correlation of assay and plant yields is more difficult with vertical retorts than horizontal retorts, although it is satisfactory in both cases. CHANNING WILSON

Midwest coals in gas-producer practice. O. L. SCALES, *Blast Furnace Steel Plant* 19, 252-4 (1931) —While there is no advantage in the use of midwestern coals from the standpoints of operation, of more satisfactory railroad service or of improvement in the product, good midwestern coal is considered, at present prices, more economical for gas producers than eastern coal. H. C. PARISH

Variations in the evolution of gases on preheating coals having reached different stages of evolution. M. LEGRAYE *Chimie & Industrie* 25, 18-21 (1931), cf. C. A. 24, 710 —Twelve coals (all but 1 from the Liège coal fields), selected so as to have comparable contents of vitrain and durain, were treated under exactly the same conditions by the same operator, as follows: volatile matter was detd as usual, a 2nd portion was heated 30 min at 400° , allowed to cool, and volatile matter was then detd in the usual way. Similar tests were carried out on 4 samples of vitrain which were sepd from the coals and which contained 10.2%, 12, 15 and 29% volatile matter, resp. Two samples of fusain, with 8.5 and 17% volatile matter, resp., both gave approx 40% volatile after preheating. The facts that the curve for the whole coal crosses the volatile-content axis at about 15-8% volatile matter, which is approx the crit volatile content sepg coking from non-coking coals that the general trend of the vitrain curve is similar to that of the whole coal but the total evolution of gases is considerably smaller, and that the preheating treatment completely destroys the coking properties of coal, suggest that a study of the chem changes undergone by vitrain carried out on carefully selected coals representing the various stages of the evolution of coal might yield valuable information on its role in the coking of coal. A. PAPINEAU COUTURE

Protection of working people and of the neighborhood in the firing of pulverized coal. W. WEITFELD *Zentr. Gewerbehyg. Unfallverhütung* 17, 142-7 (1930) —Four dangers exist: dust, poisonous gases, explosions and fire. Inhalation of dust is less dangerous than CO and other gases, which occur from excessive pressure due to blockage. Fat coal is the most dangerous from an explosion standpoint. Coal of 50% ash is particularly free from risk. Conditions causing fire and explosion are described. GEORGE R. GREENBANK

An investigation into the health of employees in gas-making plants. F. R. KERR, Commonwealth of Australia Dept. Health, *Bull* No 7; *Bull. Hyg* 6, 27 (1931) —Three hundred and seventy workers in gas plants were subjected to radiant heat, fumes and coal dust. They were divided into 3 groups according to Schneider's fatigue test, pulse rate and blood pressure. Only 105 fell into the last group which gave the poorest test. There seemed to be no ill effects of the work. Eight cases showed 5% or more carbohemoglobin, 1 case showed 15%, but no increase in red blood cells. K. recom-

merits appliances for administering O and CO₂ mixts. In acute cases of gassing.

GRUBER R. GREENHANK

Purification of town gas by means of oxide of iron. J. DEVERMAN *Gas J.* 193, 97-100 (1931).—The principal factors governing the use of iron oxide for the purification of gas are (1) the activity of the oxide for the absorption of H₂S, (2) the physical state of the oxide, (3) the concn. of the H₂S in the gas, (4) the velocity of the gas, (5) the time factor for chem. reaction and (6) the design and economic design of boxes. In practice iron oxide is used either in the form of a slurry by mixing with water and circulating through tower scrubbers or as a relatively dry, spongy mass standing on grids in covered purifier boxes through which the gas is passed. The dry method is the most commonly used. The suitability of various forms of iron oxide is discussed. When the dry method is used the oxide is usually mixed with sawdust, shavings, grain husks, cut-up straw, tan bark or oil oxide. Details are given on the operation of an iron oxide purifier at Mortlake (Australia).

P. H. HARRISON

The utilization of butane-air gas in domestic appliances. G. L. BRENNAN AND L. H. WEAVER *Gas Age Record* 67, 37-41, 40 (1931).

LESLIE H. BRAGO

The tetralin process (for avoiding naphthalene stoppages in gas mains). GEORGE WASSERBURGER *Gas u. Wasserfach* 74, 154-6 (1931), cf. *C. A.* 24, 5454, 1978.—Two methods for introducing tetralin into gas mains are given: (1) by electric heating to vaporize the tetralin and (2) cold atomization. These methods are illustrated and discussed. The formula for the amt. of tetralin to be vaporized into the main per day is $x = 5m(A + 8)/100000$, where x is the kg. of tetralin per day, m is the vol. of gas in cu. m. per day, and A is the naphthalene content of the gas in g. per 100 cu. m. Users of tetralin are warned as to its inflammable nature. Precautions are given.

R. W. RYAN

Distant pressure control (in the distribution of gas). H. C. WILKINSON *Gas J.* 193, 91-7 (1931).—The object of a distant pressure indicator is to furnish at the works a continuous indication of the pressure prevailing at some selected point on a distant district. W. describes an instrument of his own design for this purpose. It is of the proportionate current type, i. e., it maintains in its circuits a current of electricity which at all times is directly proportional to the pressure prevailing at the distant point. It consists of 2 parts: (a) the transmitter located at some distant key point, and (b) the indicator which is installed at the works. A sketch is given which shows in detail the construction of the app.

P. H. HARRISON

Steam generation in gas works practice. W. GERRARD *Gas J.* 193, 215 (1931).—It is general practice to provide for all gas works power requirements with steam. The av. gas works requires 100-1200 lb. of steam per ton of coal carbonized, for power and process requirements. Steam can be generated in 2 ways: (1) by utilizing surplus heat from the retorts, (2) by firing boilers with coke breeze. The first method is most commonly used in modern works. Any well run gas works can be self contained in "waste heat" steam alone for ordinary requirements. Coke breeze should only be used on the works when special steam requirements demand an excess over nominal figures.

P. H. HARRISON

Cheaper gas with off-peak electric power. PAUL MCMICHAEL *Gas Age Record* 67, 1-5, 10 (1931).—A central station can reduce the cost of generating electric energy to suitably steady load by operating continuously at a rate approaching that of its efficiency and selling any surplus energy for the cost of the fuel required for its generation. This surplus energy may be used to produce H and O for chem. and metallurgical processes. Pending a development of demand for this H and O the gases can be used by a gas company whenever the cost of off-peak power is no more than 1.00 mill per kw. hr. a c. The O may be used to manufacture oil gas with which to enrich blue gas even when the off-peak energy cost exceeds 2.50 mill per kw. hr. a c. It might be used for some process such as NH₃ synthesis pending development of a demand by chem. or metallurgical industries.

LESLIE H. BRAGO

Italian natural gas. M. G. IRVI AND C. PADOVANI *Atti III congresso sci. chim. pura applicata* 1930, 698-717.—The sources which contain definite, appreciable quantities of hydrocarbons lighter than CH₄ are not more than 20, all in mountains or hills. It is probable that all, or almost all, of the Apennine sources contain small quantities of higher hydrocarbons. There is no great difference in the compn. of gas from mountains and of gas from plains, except that the latter contains more CO and N₂.

P. M. SYMMES

Cleaning stack gases. JOHN B. C. KERRISAW, *Elec. World* 97, 542-6 (1931).—London power stations are washing the plant gases to eliminate dust and noxious gases. Three essentials for success of the treatment of flue gas are: (1) the placing of sprays

along the first flue to increase the humidity of the gases, with surfaces of Fe to assist in converting SO_2 to SO_3 , (2) the provision of further sprays and of wetted contact surfaces in the second flue to continue the conversion and to effect the soln. of SO_3 and its elimination from the waste gases, and (3) the provision of further surfaces moistened with alkali for the total elimination of S from the gas. The Battersea power station is planned to fulfil 14 conditions named. The plan permits the use of an overall figure of 25 tons of water per ton of coal, and the time of contact allowed for ranges from a min. of 23 secs. to a max. of 44 secs. Corrugated plates of (a) Staybrite steel (b) galvanized steel, (c) a Si-Al alloy, and (d) tinned iron are being tried out to secure the necessary catalytic action. A close-grain cast iron or a S-cast iron will be used for all pipe work and an alloy of Si and Cu or of cast Fe and Monel metal will be used for the circulating pumps and other machinery in contact with the wash water, and sprays of the centrifugal type will be constructed of a Si-Cu alloy. Cost estimates for gas washing (5% per ton of coal burned) are considered high. Results of exptl. trials of gas washing at Grove Road station of the London Power Co. in 1930 are tabulated.

W. H. BORTON

Automatic seal for water-gas plants. J. H. WAINWRIGHT. *Gas World* 94, 56-7 (1931).—The differences between the ordinary method of water gas manuf. and the back run process are pointed out and the automatic seal box of the latter process is described in detail.

F. H. BERGEM

The coal water-gas plant at Plauen. MILLER. *Gas u. Wasserfach* 74, 173-S (1931).—This water gas plant uses a "high pressure" water gas generator with a revolving grate and a distn. chamber in the upper part of the generator, a carburetor for cracking the vapors of the low temp. tar and a superheater. This gas had a heating value of 370 B. t. u. per cu. ft. and a sp. gr. (air) of 0.56. The yield was 25.4 cu. ft. of gas per lb. of coal. The cycle, analysis of the gas and some operating details are given. Although the plant has only been in operation a short time, it appears to have been successful.

R. W. RYAN

Graphical determination of the calorific value of coal gas. A. ALTON. *Gas Eng.* 43, 41-2 (1931).—The construction of nomographs for the detn. of the calorific value of coal gas is described.

CRANNING WILSON

Distribution experiences with dry coal gas. J. W. HARRISON. *Gas Age-Record* 67, 121-4, 130 (1931).—Troubles from dust, varying gravity, naphthalene stoppages and gum formation, following a change from said. water gas to partially said. coke-oven gas with water gas addn. for peak loads, are discussed. Steam injection and oil fogging have been tried with partial success.

LESLIE B. BRACG

Nature and properties of certain hydrocarbons in coal gas and their effect on meter leathers. R. S. ANDREWS. *Gas J.* 193, 155-59, 212-3 (1931).—Unsatd. hydrocarbons occurring in the light oil content of carbureted water gas, particularly styrene and indene, polymerize in the presence of O_2 to gums and resins and deposit on the meter leathers and moving parts with detrimental results. These unsatd. hydrocarbons owe their origin to imperfect cracking due to the low temp. in the carburetor. Analytical data are given on the compn. of vertical retort gas, horizontal retort gas and carbureted water gas and also on light oils scrubbed from these gases. The tests show that of the 3 gases tested horizontal retort gas is by far the least damaging to meters and carbureted water gas the most damaging. The horizontal retort gas is the lowest in light oil and this light oil is the lowest in unsatd. hydrocarbons. Meter troubles have been accentuated by the presence of large percentages of continuous vertical retort gas. Methods of preventing meter troubles are, (a) efficient methods of washing and condensation, (b) avoiding low temp. in retorts, (c) maintaining low content of O in gas, (d) addn. of blue water gas, (e) introduction of semi-chrome-tanned leather diaphragms, (f) correct selection of leather dressing, and (g) selection of a suitable position (where there is a min. temp. variation) for the consumers meter.

F. H. BERGEM

Determination of tar fog in coal gas with glass filters. HORST BRÜCKNER. *Gas u. Wasserfach* 74, 143-4 (1931).—A very effective analytical filter for detg. tar fog in gas is illustrated. The filter consists of 2 fritted glass filters in a large tube, constructed at both ends. The inlet may be partly filled with cotton to remove larger particles and the other end of the filter tube is connected to the gas meter. Most of the tar fog is caught by the cotton and the first glass filter. The gas flow is stopped when a brown coloration first appears on the second fritted filter. Directions are given for carrying out accurate tar fog detns.

R. W. RYAN

The determination of the heating value of coke-oven gas with Junker's gas calorimeter. HUBERT GAWE. *Arch. Eisenhüttenw.* 4, 75-85 (1930).—A gas calorimeter with attachments for air moistening and gas-cooling is described.

H. S. VAN KLOOSTER

vertical retort in the vertical retort and in the coke oven are, resp., benzene 61.2, 32.5, toluene 14.0, 14.0, 15.0, xylene (and light naphtha) 9.7, 21.4, 10.0, heavy naphtha 8.0, paraffins 1.6, 15.3 — loss on washing 7.1, 13.4, 5.3. The compn of the benzene varies with the time of carbonization when only one system of carbonization is used. A silica regenerative coke oven battery gave crude benzene of the following compn when operated under the carbonizing times of 29.5 hrs and 10.2 hrs, resp.: 29.5 42.9, 62.0, toluene 20.4, 17.9, solvent naphtha 6.8, 5.3, paraffins up to 140° 11.1, C₂₀ 0.1, 0.1, washing loss 8.5, 1.1 and benzene residues 7.1, 5.2. Other constant factors besides time which influence the compn of crude benzene are (1) the size of oven, (2) carbonizing temp., (3) uniformity of heating, (4) the density to which the oven is filled with the charge, (5) class of coal, (6) moisture in the coal and (7) operation of the oven. Figures are given showing the absorption efficiencies of benzene in gas oil at varying temps. F. H. BERGMAN

The refining of motor benzene by means of silica gel. R. C. GAUDIN. *Gas World* 94, C (King Section 9-12(1931)). The silica gel process gives a greatly increased yield over the usual H₂SO₄ refining process. Where 72% is the normal yield in the old type of plant employing a once running still for refractionating the crude recovered benzene silica gel gives 85%. In a modern plant with efficient stripping stills, and the primary once running still eliminated, where 85% is being obtained, the yield with the silica-gel process is 91%. The following data give a comparison of the results of the H₂SO₄ and the silica gel processes. By starting in each case with crude benzene of the following % compn: diolefins 1, olefins 10 and aromatics 89 the resultant benzene refined with strong H₂SO₄ is diolefins 0, olefins 0 and aromatics 89 parts. With dil H₂SO₄ and with silica gel the resp. figures are 0.25, 2.5, 89 and 0, 0.5 and 89. The high yield in the silica gel process is due to the retention of the stable unsatd compds which are not only not harmful to motor benzene but also have a definite antiknock value. In the silica gel process the polymerization of the undesirable unsatd compds is started by the treatment with dil H₂SO₄. The polymerization is accelerated by the catalytic action of silica gel at 150°. A unit wt of silica gel will treat 1500 times its wt of oil before activation becomes necessary. The general layout of a silica gel refining plant is given. F. H. BERGMAN

Construction of waterless gas holders. L. HARTLEY. *Gas World* 94, 61-2(1931). — A review is given of the structural features of waterless gas holders. F. H. BERGMAN

Measurement of the amount of dust and tar in generator gas. GUSTAV NEUMANN AND FRANZ STRAUTNER. *Arch. Eisenhüttenw.* 4, 151-4(1930). — An app. was developed for the detn. of dust and tar in generator gas, which is suitable for the collection of the coarse and fine dust coming from the generators over longer periods of time. The 2 main parts are the collection chambers, also serving as cooling chambers, and the filters. A tube of NCT steel proved suitable as a collection tube. Cotton was used as a filter material. An av. analysis of the gas showed the presence of 8.65 g. of dust, of 11.35 g. of tar and of 55.80 g. per cu. m. of dry gas. LEOPOLD PESSEL

The free carbon in coal tar. J. MARCUSSE. *Chem.-Ztg.* 54, 795-6(1930). —Tars from a vertical still and from a horizontal still are compared. The amt. insol. in CCl₄ was 7 and 24%, resp. The benzene sol. part contained 8.6 and 0.5% of hydroxy acid, 7.3 and 16.3% of pyridine-sol. resins, 18.4 and 52% of pyridine insol. resins and 0 and 51.2% coke, resp. A. L. HENVE

Extraction of bases from primary tar by water of condensation. A. TRAVERS AND FRANQUIN. *Compt. rend.* 191, 951-2(1930). —Detn. of the amt. of pyridine bases in the water of condensation from tar is difficult because their concn. is only about 0.1% and NH₃ is present in 6-7 times as high a concn. Extn. of the bases by an org. solvent, e. g., PhH, takes up about 90% of the phenols present in the water and a considerable proportion of the bases but complete extn. is impossible because an equil. is established between the aq. and the hydrocarbon phases. The presence of the pyridine bases improves the extn. of phenols on account of the formation of addn. compds. The best lab. procedure is to sep. the phenols by distg. from 30% NaOH, wash the distillate which contains all the NH₃ and bases 5 times with 50% of its vol. of Et₂O, distil off the Et₂O at 0°, remove the NH₃ from the residue by blowing CO₂-free air through at 0°, then add 40% NaOH soln. and read volumetrically the amt. of base liberated. Applied to the PhH used in extg. phenols, this method shows 7-8 g. of bases per L. of PhH. C. H. PEEB

Determination of phenols in primary tars. I. UBALDINI AND O. MOCHI. *Atti. III congresso naz. chim. pura applicata* 1930, 782-91. —Detn. of phenols in tars may give results varying from 30 to 20% phenols. An improved method is to put 20 g. of tar in a vacuum walled vessel, add 10 cc. of NaOH soln., d. 1.332, the tar and NaOH soln.

being at the same temp before mixing. The 2 are mixed, stirred with a glass rod and the max temp reached during the reaction is observed with a thermometer reading to 0.1°. The max temp is reached a very few sec after mixing and lasts long enough for reading. Dividing the rise in temp by 0.12 gives directly the % of phenols present.

I. M. SYMMES

Production of hydrocarbons from phenolic tars. C. L'ADOVANI AND L. DE TARTAGLIA. *Atti III congresso naz chim pura applicata* 1930, 792-810.—It was possible to convert into hydrocarbons the primary lignite tar fraction passing over at 200-300° and sol in boiling water sol by heating to 750-800° in an atm of H₂ and nascent H formed by the action of water in Fe. The best results were obtained with 100 g of phenols heated for 2 hrs at 780° in a tube 1 m long in the presence of water vapor and H₂ forming 35-40 g of dephosphated oil b below 200°, about half of which was pure C₁₁. This yield can be increased by increasing the size of the tube and thus the contact surface.

I. M. SYMMES

Investigations on commercial primary lignite tar from Valdarno. M. G. L'UCCI, C. L'ADOVANI AND L. TARTAGLIA. *Atti III congresso naz chim pura applicata* 1930, 767-77. Total analyses and fractional distn of lignite tars are shown. I. M. SYMMES

Coking and decomposition heats of coal. H. LUNN, F. FERRE AND KARI VOGTHER. *Gas u Wasserfach* 74, 98-101, 122-8, 118-54, 178-82 (1931), cf C A 22, 3017. A new method has been devised for detg heats of coking of coal with greater accuracy than heretofore by altering the method of operating the calorimeter. This new method is based on the fact that the heat content of the empty system is a straight line function of temp between 100° and 1100° and is independent of the heating time or amt of heat applied. The heat of coking is the difference between the total amt of heat supplied to the system until the heat taken by the calorimeter up to the shut off point, less the heat content of the system at the given temp. The heat content of the gaseous products need be known only for calg the heat of decompn. Heat of coking curves are shown for 15 coals over the range from 100° to 1200°. These curves vary widely with the various coals, and no relation can be found between the curves and analytical data for the coal. The heat of coking of a given coal is reduced by storage and this decrease is somewhat proportional to time of storage. Heats of coking and heats of decompn of 30 coals at 1000° are plotted against each other and fall on a smooth curve, permitting the estn of the heat of decompn of a coal when the heat of coking is known. The max heat of decompn of any coal is +50 kg cal/kg while many coals show neg heats of decompn. The heat of coking of 'pure' coal was shown to be 3.55 kg cal per kg. This investigation is being continued in the hope of applying the results to the coking process.

R. W. RYAN

Coking a banded bituminous coal. C. P. LUNN. *Iron Coal Trades Rev* 121, 910 (1930); 122, 3, *Gas World* 94, Coking Section, 2-6 (1931). See C A 25, 1650.

LESLIE H. BRAGG

Coke-oven operation and maintenance. F. J. PRINKE. *Am Gas J* 134, No. 2, 41-4, No. 3, 37-41, No. 4, 75-8 (1931), cf C A 25, 1761.

F. H.

The formation of fractures in coke. R. G. DAVIS AND R. V. WHELFER. *Gas World* 94, 16-7 (1931).—Carbonization results confirm Flinn's previous work that cracks appear near the walls of the oven shortly after the start of the coking process rather than at a later stage, a short time before the carbonization is completed.

I. H. BERGMAN

Coke for domestic fires. H. J. HIGGINS. *Gas World* 94, Coking Section, 23-4 (1931).—A review is given of the progress of coking technology.

P. H. BERGMAN

Continuous verticals. W. A. CURRIE. *Gas J* 193, 229-31, *Gas World* 94-5 (1931).—The effects of temp, rate of carbonization and coal size on the products of carbonization are discussed. A high temp is favorable to a large throughput, and gives a high yield in thermal per retort. A formula is given for the calen of the optimum throughput of coal based on the free space in the charge and the swelling power of the coal.

I. H. BERGMAN

Circulating-stream coke ovens. O. PRINKE. *Stahl u Eisen* 50, 761-7 (1930).—This type of coke oven has twin flues, and a portion of the burnt gas is drawn by suction back into the incoming stream of heating gas and air so as to retard the combustion, avoid overheating and regulate the heat flow from gas to coking retort. A furnace of this kind is described and illustrated diagrammatically, and a brief account is given of its operation.

B. C. A.

Some additional volume data for superheated steam (SMITH, KRYSS) 2. The relative merits of gas, oil and electricity for industrial purposes (HOLKINSON) 13. Miner-

alogy of the coal fields (BELLIERE) 8 Automatic analyzers of gas mixtures based upon thermal conductivity of gases and their industrial uses (IT PROBERVILLE) 1 Lignin, humic acid and humin (LICH) 10 Utilization of sewage gas (KUTSRL) 14 Chemical properties and examination of tar and pitches (BOCKSHAMER) 22 Automatic gas-analysis apparatus (Brit pat 337,842) 1 Recovery of oils from mixtures such as destructive hydrogenation products of brown coal (U S pat 1,794,845) 2 Freeing tars from suspended solids (Brit pat 338,154) 22 Destructive hydrogenation (Brit pats 337,671 and 338,150) 22 Fuel from petroleum distillation residue (U S pat 1,793,014) 22

Von den Kohlen und den Mineralölen, 1930 Bd. III. Ein Jahrbuch für Chemie und Technik des Brennstoffe und Mineralöle Edited by Fachgruppe für Brennstoff- und Mineralölchemie der Vereinus Deutscher Chemiker Berlin Verlag Chemie 238 pp M 17

Fuel, etc. TROCKENUNGS, VERSCHWELUNGS, UND VERGASUNGS G M B H Ger 482,559, Feb 3, 1928 Material to be treated thermally in dust form, e g, coal dust, is prevented from mixing with gaseous products by elec cohesion

Fuel briquets DONALD S ANDREWS U S 1,793,833, Feb 24 In making fuel briquets with an anthracite base, culm 85-90, asphalt 5-10 and pulverized bituminous coal about 5% are used together The asphalt is rendered freely fluent by heating, the culm is heated to about the same temp and mixed with the asphalt and the bituminous coal is then added and intermixed

Fuel briquets HANS HERTIG SWISS 142,960, April 17, 1930 Briquets are made by mixing sawdust with 8-30% of adhesive substance (e g, water glass soln) and pressed

Pressing fuel briquets. FIRMA CARL STILL Ger 514,888, Sept 3, 1929

Carbonizing fuel briquets MAUREL INVESTMENT CORP Brit 337,600, Oct 31, 1929 Briquets of ground coal and a hydrocarbon binder are passed through a liquid seal into a distg zone where they are heated solely by indirect heat, and then through a second liquid seal and into a hardening zone and finally into a vertical retort. Various details of app and operation are described Cf C A 25, 102

Distilling and gasifying solid fuels. METALLGES A G Brit 337,721, May 2, 1929 Gasification and distn of solid fuel are effected as it passes through a vertical shaft comprising a drying zone, distg zone, gasifying zone and cooling zone Steam is superheated by heat from preheated refractory bodies and is introduced below the gasifying zone Gases for heating the fuel are passed through the app countercurrent wise and tar or heavy or light oils may be mixed with the gases to increase the heating value of the gases produced, or O may be added to the gas or steam supplied to the app Various details of construction and operation of the app are described

Catalytic treatment of volatile engine fuels FERNAND C F PORTAIL (to Soc anon le carbone) U S 1,795,037, March 3 Air and fuel are brought into contact with very active porous C at a temp of 300-400°, in order to facilitate subsequent complete combustion with addnl air App is described

Destructive hydrogenation N V DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ Brit 338,192, Aug 14, 1929 Hydrogenation under pressure of such materials as tar, petroleum products and residues, lignite and cellulose is effected with a catalyst consisting of a Mo compd carried on adsorbent material such as adsorption charcoal or finely divided coal In prep the catalyst, an aq soln of NH₄ molybdate, acidified with H₂SO₄, may be treated with H₂S to form a blue colloidal soln to which "Carboraffin" or finely divided brown coal is then added, followed by filtration, washing and drying first at ordinary temp and then at 50-60° Various examples with details of the hydrogenation of brown coal and gas oil are given Cf C A 25, 1362

Fuels from coal and oil TRENT PROCESS CORP Brit 337,920, Oct 20, 1928 In a process involving coal and oil amalgamation, as generally described in Brit 228,802 (C A 19, 3011) and Brit 262,302 (C A 21, 3733), the amalgamation is first effected with a light oil such as "navy fuel oil" and the process then completed by addn of a heavy oil (suitably with intermediate removal of ash content and change of water) Proportions used may be coal 85, light oil 5.0-7.5 and heavy oil 10-7.5% Cf C A 24, 1735

Decomposing coal DEUTSCHE BERGIN A G FÜR KOHLE- UND ERDÖL-CHEMIE (Wolf Grote, inventor) Ger 517,317, July 12, 1928 In the destructive hydrogenation of coal at a high temp and pressure, the production of coke like and asphaltic products is hindered by addn of a mixt of Fe₂O₃ 75-60 with Na₂CO₃ 25-40 parts, with or with-

out a little alk earth carbonate About 10 parts of the mixt may be added to 140 parts of coal

Apparatus for pneumatic separation of solids of different densities as in purifying coal. J F C FRIEND Brit 317,458, Aug 2, 1929 Structural features

Furnace for coking coal and heating a boiler or the like by the heat developed in coking. T KOMATSU Brit 338,435, Feb 6, 1929 Structural features

Retort for destructive distillation of coal or lignite. H HARDY Brit. 338,047, Jan 3, 1929 A drum is provided with a hollow shaft for admission of heating gases and evacuation of combustion products The drum contains combustion chambers, distn chambers and gas outlet chambers, and progressive heating is provided either by rotation of the retort on the hollow shaft or by rotation of the latter while the retort is stationary Various structural details are described

Plate drier for lignite 'EINTRACHT' BRAUNKOHLENWERKE UND BRIKETTFABRIKEN and MAX MAYER Ger 514,711, Feb 7, 1929

Smoke consumers GEORG KORNECK Ger 514,975, April 15, 1926 Addn to 500,932 (C A 24, 4670) App for burning the smoke and flue gases in furnace plant is described

Condenser suitable for steam condensation. PELLE ANDERSSON (to Allis Chalmers Mfg Co) U S 1,792,796, Feb 17

Combustible gas containing hydrogen and carbon monoxide. HERBERT A HUMPIREY (to Atmospheric Nitrogen Corp) U S 1,794,232, Feb 24 In a continuous-gas-producing process a mixt comprising steam and O, preheated to above 1000°, is passed into and through a gasification chamber and finely divided solid fuel is fed into the chamber to contact and react with the preheated mixt and to generate gas at such a high temp that CO and H substantially free from CH₄ or other hydrocarbons and CO₂ are produced, the hot generated gas is led into and through a second preheating stage and from the latter combustible gas is withdrawn A reversal of flow of the mixt contg steam and O is effected whenever its temp immediately prior to its contact with the fuel tends to fall materially below 1000° and (irrespective of the direction of gas flow) the particles of solid fuel are maintained in suspension in a relatively large mass of enveloping gas in the reaction chamber App is described

Gas from coal distillation. I G FARBENIND A G Brit 338,153, May 14, 1929. Bituminous coal is distd by heating it in granular or finely divided form in a heat accumulator which is periodically heated to about 900°. residual coke may be gasified and the producer or water gas passed through the heat accumulator to heat it Various details of app and operation are described

Dry purifier for coal gas, etc FRANCKE WERKE A G Ger 517,320, Aug 24, 1928 Structural features are described

Fuel gas. MATHILDE LOTZ NÉE BLICKLE, AUGUSTE LOTZ, ARTHUR LOTZ, MATHILDE SPANKE NÉE LOTZ, MARIA KOWSKY NÉE LOTZ, ROBERT LOTZ and WALTER LOTZ Ger 517,414, Apr 22, 1928 Addn to 513,233 (C A 25, 1364) When using coal that shrinks only slightly or not at all on gasification, the method of Ger 513,233 is modified by gasifying the coal so that a layer of partly gasified coal is obtained on a layer of coke Steam is then led into the coke layer, the water gas so produced passing through the upper layer

Gas mixtures. GIULIO NATTA SWISS 141,837, Mar 7, 1929 A gas mixt of CO, H and CO₂ contg at least twice as much H as oxide of C, is obtained by passing an O and steam current (contg at least twice as much steam as O) over C at about 700° and not over 750° Cf C A 24, 5765

Mixed water gas and coal gas EDWARD A DIETERLE U S 1,792,632, Feb 17 A hot bed of solid carbonaceous fuel is intermittently blasted with air, and steam is intermittently passed through it upwardly Powd carbonaceous material is intermittently introduced above the hot fuel by causing the powd material to pass downwardly through a vertically disposed retort positioned above the fuel bed and externally heated by the rising hot gases which are generated to effect distn of the powd material before it encounters the hot fuel and the generated blue water gas, the resulting gases and entrained vapors are subjected to further fixing by passing them through vertically arranged elongated heated passages above the main fuel bed and around the retort App is described Cf C A 24, 949

Water gas I G FARBENIND A G (Fritz Winkler and Rudolf Augsten, inventors) Ger 517,469, Feb 22, 1928 Addn to 437,970 For the continuous manuf of water gas from granular fuel, which may contain dust, steam is blown into a shaft contg the fuel so that the fuel is whirled about in the shaft, and the necessary heat is

Continuous water-gas production. I G FARBER and A G Brit 337,807, Aug 1930. Fuel such as brown coal, small coal, semi-coke or peat is passed downwardly through externally heated vertical retorts supplied with steam (preferably superheated to at 500° while the fuel is heated to about 500° and the retort heating flues are heated to about 1200°). Distn. gas from the upper part of the retort is collected separately from the water gas from the lower portions. Various details for continuous production are described also of app. construction.

Automatic control system for electric valve operation in apparatus for generating water gas. JAMES KENNEDY U. S. 1,794,906, March 3. Mech. and elec. features.

Apparatus for producing oil gas by partial combustion of heavy oil with superheated air and steam. C. CHILLOWKY Brit 338,281, Jan 18, 1929. An atomized fuel, superheated steam and air issues from a nozzle and impinges against a crucible which may be of graphite or carbonundum or non-oxidizing steel and in which eddy current are set up; the flames pass in contact with the outer walls of the crucible and the products of partial combustion are then discharged through a catalyst (or the crucible itself may constitute a catalyst). Various details of construction are described. Cf. C. A. 24, 283.

Gas producer. HARRY F. SMITH (to Gas Research Co.) U. S. 1,794,478, March 3.

Automatic control is provided for effecting communication between the off-take and a vent pipe or delivery main in accord with the quality of the generated gas, and various structural details are described.

Gas-producer plant. POWER GAS CORP., LTD., and N. E. RAMBUSH. Brit. 337,425, Aug 9, 1929. Various details are given of the construction and operation of a plant comprising a generator or producer, a waste heat boiler, a generator jacket boiler and a tubular superheater, two sets being used (alternately on "run" and "blow") for making water gas.

Gas producer (with a rotatable fuel shell and rotatable ash pan). ANSON K. BRADLEY (to Morgan Construction Co.) U. S. 1,793,162, March 3.

Gas producer (with a rotary ash pan). PHILIP S. HULT (to Morgan Construction Co.) U. S. 1,793,618, Feb. 24. Structural features.

Gas producer adapted for directly heating steam boilers. W. B. CHAPMAN. Brit. 337,388, July 29, 1929.

Apparatus for distributing granular fuel to gas producers or for similar purposes. A. JADOTT. Brit. 338,141, July 11, 1928. Structural features.

Domestic gas plant for generating gas from oil and vegetable materials. JOSEPH S. BIRCHER (one fourth each to E. A. Rodgers, John S. Fogarty, Gerald Redmond and Homer L. Baughman) U. S. 1,795,005, March 3. Structural features.

Apparatus for washing flue gases with water sprays. J. T. BARON and J. B. CLARKE. Brit. 337,426, July 1, 1929. Structural features.

Gas holder. E. CHURCH. Brit. 337,711, March 7, 1929.

Purifying furnace gases. SIEMENS-SCHUCKERTWERKE A-G. Brit. 337,448, Jan. 9, 1929. Furnace gases which are to be passed through a separator such as an electrostatic pptn. app. are preliminarily brought to a uniform temp. and humidity by an app. automatically controlled by a thermometer and hygrometer. Various details of app. are described.

Removing naphthalene from coal gas. C. OTTO & CO., GEB. (to N. V. Sijka en Ovenbouw Mij) Brit. 337,723, May 17, 1929. A solvent for naphthalene such as xylene is introduced into the gas in atomized or nebulized form, followed by electrostatic pptn., sepn. of solvent from the naphthalene and reuse of the solvent.

Separating ammonia and hydrogen sulfide from gases. CHRISTIAN HANSEN (to I. G. Farbenindustrie A-G) U. S. 1,793,121, March 3. In treating gases such as coal gases most of the NH_3 content is sepd. by treatment with water, and the residual gases are then washed in 2 stages, in the first stage, with the aid of a soln. contg. NH_4 thio-sulfate, SO_2 and NH_4 polythionate, such a part of the H_2S is pptd. in the form of S that for the washing of the residual H_2S in the second stage a neutral NH_4 sulfite-bisulfite soln. suffices which is produced from the sepd. NH_3 and S.

Ammonia and hydrogen sulfide absorption simultaneously from industrial gases. CHRISTIAN HANSEN (to I. G. Farbenindustrie A-G) U. S. 1,793,120, March 3. Absorption of NH_3 and H_2S from gases such as coke-oven gas or illuminating gas is effected in a 2 stage process in the first stage of which is used "ammonium sulfite disulfite" wash liquor having a lower ratio figure of SO_2 to NH_3 than about 1.5, and in the second stage of which there is used an "ammonium sulfite bisulfite" wash liquor having a higher ratio figure of SO_2 to NH_3 than about 1.5.

Storing explosive gases. AUTOGEN GASACCUMULATOR A-G. Ger. 514,722, Mar. 2,

1028 Porous raw material is extd and evacuated to remove substances such as resin, fat, gas, etc., and then filled into vessels to store such gases as C_2H_2 . Cf. C. A. 24, 4021.

Storing acetylene, etc. INDUSTRIES A-G ZWERNIGERLASSUNG WAGRO-DISSOUGASWERKE (Gustav Offe, inventor). Ger 517,351, Sept 24, 1929 Peat or like bituminous substances mixed with active C, or with flue dust from boiler furnaces, are used as a filler for containers in which C_2H_2 or other gas is to be stored in soln under pressure

Distilling tar. T O WILTON and CHEMICAL ENGINEERING & WILTON'S PATENT FURNACE CO., LTD Brit 337,581, June 13, 1929 In distg tar by a process such as that described in Brit 307,577 (C A 23, 5308), more intimate mixt of the crude tar and the pitch from the process is obtained by introducing them side by side through pipes into a const level tank Various details of app and operation are described

Coking tar or molten pitch, etc. E O RHODES (to American Tar Products Co.) Brit 337,800, Feb 9, 1929 The material to be coked is sprayed into a heated coke oven in the presence of an inert gas such as steam (preferably after a preliminary distn of tar if the latter is used) Various details of app and procedure are described

Coke and distillation products from coal. RICHARD H CAIR and CORNELIUS B. WATSON (to Pure Oil Co.) U S 1,793,838, Feb 24 Various details of app and procedure are described for low temp distn

Vertical-flue coke oven. C OTTE & Co., GES (to N-V Silica en Ovenbouw Mij) Brit 337,667, May 9, 1929

Vertical-flue coke oven J VAN ACKEREN (to Koppers Co.) Brit 337,801, April 25, 1929 Numerous details of construction and operation are described

Vertical coking retort and charging and discharging apparatus T TOPPING and J. S BLACK Brit 337,842, March 27, 1930

Coke-oven gas-reversing valve COLLIN & Co Brit 338,040, Dec 27, 1928.

Coke oven (door construction). CARL STILL U S 1,795,239, March 3

Apparatus for operating a coke-oven door. RAYNARD CHRISTIANSON (to Koppers Co.) U S 1,793,458, Feb 24 Structural features

22—PETROLEUM, LUBRICANTS, ASPHALT AND WOOD PRODUCTS

W. F. FARAGHER

Quinquennial review of the mineral production of India for the years 1924 to 1928. Petroleum. E H. PASCOE *Records Geol Survey India* 64, 257-73(1930); cf. C. A. 24, 4245—Indian oil production from 1924 to 1928 declined, although 1928 was a banner year The country contributes less than 1% of the world's marketed supply. World's production and consumption are reviewed and the individual Indian fields discussed in detail

ALDEN H. EMERY

The total heat and specific heat of a series of fractions of petroleum oils and their relation to other properties. H R. LANG, R. JESSEL AND A H. STEED. *J. Inst Petroleum Tech* 16, 783-813(1930)—This paper gives the 1st results of a planned systematic study of the phys and particularly the thermodynamic properties of petroleum oils A Mini (Boroeo) petroleum was sepd into 5 fractions with little overlapping. Sp heat, total heat, their variation with temp, n , mean mol wt, sp. gr and coeff. of expansion were measured Variations of the sp heat, temp curve from the straight-line relation are explained at low temps by the presence of C_6H_6 with its high f. p., at high temps. by the presence of vaporized mols and intermediately, by some aggregate forming tendency of the mols Total heat above 100° was measured by Callendar's continuous-mixt method. Sp. gr. and n were strictly linear with respect to mean mol wt.

EMMA E. CRANDAL

The isolation of the isomers of hexane from petroleum. JOHANNES H. BRUUN AND MILDRED M. HICKS-BRUUN *Bur Standards J. Research* 5, 933-42(1930)— n - C_6H_{14} , Me_2CHCH_2Me , $Me_2CHCH_2CH_2Me$ and $MeCH_2CHMeCH_2Me$ were isolated from an Oklahoma crude petroleum, but the 5th isomer, Me_2CCH_2Me , was not found. Ordinary fractional distn concd the hexanes in a set of const-boiling mixts, the other constituents of which were ring compds The const-boiling mixts were broken up by distn after the addn of an alc (MeOH or EtOH) which was later removed from the distillate by washing with H_2O . The f. ps for air-satd $Me_2CHCH_2CH_2Me$ and $MeCH_2CHMeCH_2Me$ are $-143 \pm 0.5^\circ$ and $-118 \pm 0.5^\circ$, resp The eutectic for the system CO_2 - $Me_2CHCH_2CH_2Me$ is $-153 \pm 0.5^\circ$.

R. E. SCHAAD

The electrical conductance of hydrocarbons in thin films. L. BACNINGHAUS *Compt rend* 192, 151-3 (1931) cf C A 24, 4142—Films of oil 1 μ thick become highly conductive to 110 v d c and remain so until the electrodes are sepd by 15 μ . Conductance is always preceded by a semi-conducting stage. This is doubtless due to ions similar to those in gases. V F HARRINGTON

Refining mineral oils with adsorbents. BALDISSARAE SALADINI *Atti III congresso naz chim pura applicata* 1930, 581-607 cf C A 24, 1963—Tests were made with shale oil but the results apply to other high S oils. Active C decolorizes indifferently and desulfurizes little. Decolorizing earths give rather complete decolorization and appreciable desulfurization. SiO₂ gel decolorizes less than the above, but desulfurizes better. In general desulfurization follows the adsorption of colored substances. Bauxite calcined at high temp. has an appreciable desulfurizing effect and polymerizing properties. Alternate use of bauxite and decolorizing earths on crude oil gives the best results and an economy in adsorbent. Decolorization is rather complete, desulfurization reaches 50%, the oil has a pleasant, ethereal odor and is stable. Regeneration of the treating materials is effected by calcination, which allows recovery of oil and S. I M SYMURA

Separation of obstinate crude-oil emulsions. The L. Berkhan patent. F. PINTAL *Erdol u. Teer* 6, 475-80 (1930) Caspian crude oils, emulsified with H₂O that contained clay and MgCl₂ could not be sepd by ordinary settling. Serious corrosion troubles made drying operations or utilization as boiler fuel impossible. The possibility of centrifugal sepn was demonstrated in lab. expts. The Berkhan method is in successful operation. Sepn is induced by an admixture of 0.5% of kerosene-naphthene acids and settling of the warm mixt. It was shown experimentally that the emulsions were stabilized by traces of ozocerote. Naphthene acids dissolve the stabilizer and lower the surface tension of the oil. K H ENGEL

Cracking processes in the Russian petroleum industry. L. SINGER *Petroleum* 2 25, 893-906 (1929), 26, 482-91 (1930) cf C A 24, 3644, 4619-20—Three types of cracking processes are described which serve for the production of (1) fuel oil of a better quality as regards η and cold test, cracking being carried out at 425° and under 10-60 atm. (2) gasoline and fuel oil, and (3) gasoline and coke, in which the treatment is carried out at 450° and 20 atm. These processes were utilized in the cracking of Gromni and Surakhani mazouta and of various heavy crude oils. The results obtained with Gromni mazout show that gentle cracking (first method) causes a decrease in the cold test of the fuel oil obtained. In the second method with a 1% yield of coke, 35% of cracked benzine is obtained. The quantity of coke formed is independent of the temp. but is const. with equal gasoline yields. With 15% of coke in the third method 57% of cracked gasoline is produced. The following conclusions have been reached: (a) Greater yields of light products cause an excessive formation of coke and deteriorate the quality of the residue. (b) The velocity of the reaction is doubled for each rise in temp. of 10°. (c) The rate of coke formation depends on the chem. compn. of the cracked petroleum products. Paraffinic products yield less coke than paraffin-free or asphaltic products. Kerosene yields practically no coke, solar and spindle oils yield very little, machine and cylinder oils yield large quantities and tars greatly increase the formation of coke. (d) Increase of pressure causes a decrease in the content of unsatd. hydrocarbons. paraffin hydrocarbons yield no coke on cracking; while aromatic hydrocarbons tend to give significant quantities. (e) The extent of the production of cracked gasoline and kerosene from heavy crude oil and mazout is practically the same. (f) The details of a cracking process are detd. not so much by the yields of cracked gasoline (which are the same for a given temp. and cracking time with heavy crude oils and mazout) as by the yields of coke, which vary for different petroleum products. The cracking of Gromni oil tar and paraffin tar is not important on account of the excessive quantities of coke formed. B C A

Refining of cracked gasolines. P M FOMONOV SCHMITZ *Erdol u. Teer* 6, 493-4, 510-2 (1930)—Raw gasoline treated cold with NaOH (30° B ϵ) or hot with fuller's earth showed discoloration and gum formation on standing even in the dark. Washing with NaOH should be followed by treatment with H₂SO₄ (strength 1%) A gasoline of excellent color and stability was obtained by a vapor phase purification (100-20° above the highest b p) over Fe₂O₃. Various brands of fuller's earth instead of Fe₂O₃ were less effective. Dry CaCl₂ and NaOH were practically ineffective. K H ENGEL

Compression pressure is controlling factor in inducing engine knock. SANDOR D. RUBENZ *Automotive Ind.* 63, 20-3 (1930)—Antiknock rating of a fuel is detd. on a single-cylinder Liberty engine by comparing the max. of the manifold vacuum-air fuel

ratio curve with those of similar curves obtained for known blends of com C₁₂H₂₂ in domestic aviation gasoline "The constancy of the manifold vacuum points of incipient knock leads to the conclusion that, other things being equal, the compression pressure is the controlling factor in inducing knock" Antiknock value of gasoline for air-craft engines can be detd. only if the base fuel is specified when C₁₂H₂₂ rating is given With "cold running" (large central electrode) and "hot running" (small central electrode) spark plugs, manifold vacua at incipient knock are almost identical In this Liberty engine the tendency to knock is greater with 2 plugs than with 1, although 2 plug ignition is usually stated to reduce knocking R E SCHAAD

X-ray studies on paraffin wax and petrolatum. SHINSUKE TANAKA and AKIRA TSUJI *Mem Coll Sci Kyoto Imp Univ. Ser A*, 13, 369 73(1930)—The effects of temp and mech operations were studied by taking photographs of the x ray diffraction patterns Influence of Temp—The photographs taken at temps lower than the m p have many spectral lines characteristic of cryst powder photographs These lines generally decrease in intensity as the temp rises and vanish when the substances are melting Amorphous bands appear in other positions One intense line, however, remains until melting is reached and then suddenly vanishes This transformation occurs in a different manner for each substance Effects of Mech Operations—The effect of compression and rolling at room temp was studied with paraffin wax with the following results (1) The wax acquires a fibrous structure, the axis of which is perpendicular to the surface of compression and rolling (2) The crystal planes which produce the most intense reflection are also nearly perpendicular to the surface of compression and rolling C J HUMPHREYS

Electrochemical oxidation of paraffin II. I A ATANASIU *Ber* 64B, 252-60 (1931), cf C A 24, 2281—Paraffin has been oxidized to acids by nascent O in the presence of catalysts The lower the m p of the paraffin, the easier the oxidation The app and conditions of the expt are described V F HARRINGTON

What determines the value of absorption oil? F L KALLAM *Chem Met Eng* 38, 78-81(1931)—The phys characteristics of 24 new oils and 34 used oils in use as absorption oils in various refineries are presented in the form of curves to show the predominating properties which det the suitability of a given oil for this purpose It is concluded that the mol wt of the oil should be of the order of 150 to 200 The av b p of a proposed absorbent should be specified because it is the index of the initial b p, the vapor pressure and the gravity Max distn residuum or the demulsibility factor, freedom from wax sepn at the lowest operating temp and viscosity should also be included in absorption-oil specifications D F. BROWN

Notes on the viscosity-temperature relations of lubricating oils. W L BAILLIE *J. Inst Petroleum Tech* 16, 643-60(1930)—By calcs from data in Arehbutt and Deeley's *Lubrication and Lubricants*, it was shown that Fekart's formula for the relation of temp to viscosity is as accurate for kinematic as for abs viscosity Results calcd by it agree with exptl results within about $\pm 2.5\%$ It is a special case of the general formula given by Slotte, $V = b/(t + a)^n$, in which t is temp and V is kinematic viscosity An equation was derived relating efflux time in sec in the Redwood viscometer to viscosity $AT - B/T = \pi\theta^2$ A and B are consts of the instrument, w and n , of the oil, T is time of efflux, θ is the temp Rigorous temp control is most important with viscous oils A temp variation of 0.14°F may cause a change of 1% in viscosity A table is given for conversion of Redwood seconds to kinematic viscosity EMMA E CRANDAL

Inorganic lubricants. IV. Lubricants for temperatures above and below normal. WILLIS A DOUGHTON *J Am Chem Soc* 52, 4858-60(1930), cf C A 25, 196—Aq solns of HPO_4 and NaPO_4 mixts of salts melted in their water of crystn with an inert material such as kaolin or graphite and solns of $\text{Ca}(\text{NO}_3)_2$ in melted KNO_3 are good lubricants at temps both above and below ordinary A lower limit of -75° and an upper one of 360° have been reached EMMA E CRANDAL

Amount of lubricating oil burned in the gasoline engine. CLARKE C MINTER AND WM J FRYN *Ind Eng Chem* 23, 285(1931)—The amt of oil burned in a gasoline engine was detd. by operating with H and analyzing the exhaust gases L W. T. C.

Effect of carbon black on insulating oils. W. B WIEGAND, C. R BOGGS AND D. W. KITCHEN. *Ind Eng Chem* 23, 273-6(1931)—The elec. properties of insulating oils are improved by treatment with C black L W. T. CUMMINGS

Improving petroleum asphalts. A I VOROBYOV AND N I. LOGVINOVA *J. Petroleum Ind* 1930, 439 (in Russian); *Asphalt u Teer* 33, 876-8(1930) (in German)—Still residues of low-resin and high paraffin content may be converted into satisfactory asphalts by vacuum distn, followed by the customary air blowing Asphalts of the

naphthene aromatic type, rich in desirable resins but deficient in asphaltene, should not be blown. Blending these asphalt with asphalt of the Me naphthene aromatic type is more economical and yields a technically superior product. K. H. ENGL.

Determination of asphaltene content. F. J. NELLESTEIN AND N. M. ROODENBURG. *Chem. Ztg.* 54, 519 (1930). —As a solvent for use in detg asphaltene, Li_2O is superior to normal benzene because it is a homogeneous substance not requiring standardization; its flocculating properties are combined with low surface tension, and it has a greater dissolving power for the hydroxy acids in asphalt produced by oxidation and for certain constituents of coal tar. R. F. SCHLAAD.

Separation of bitumen from bituminous sands. K. A. CLARK. *Nature* 127, 199 (1931). Cf. C. A. 24, 2873. —The bitumen sepd. by means of solns. of Na_2SiO_3 contained 1-4% of mineral matter. Difficulties to duplicate results were encountered. Improvements were obtained by giving the sand a preliminary wash with cold H_2O and then treating with alkali to bring the pH value up to 6.4. The water removes a part of the clay and reduces the necessary quantity of alkali. The acidity of the sand is considered to be due to the presence of Fe^{++} and Fe^{+++} salts carried into the beds by ground water. I. M. LEVINE.

The estimation of wax in pitch. W. LITTLEJOHN AND W. H. THOMAS. *J. Inst. Petroleum Tech.* 16, 814-24 (1930). —As the outcome of a study of what factors must be controlled in order to get results approaching reproducibility for the % of wax in pitch, a cracking method is recommended. The pitch is distd. at the rate of 40-50° temp. rise per min. to 450°. Distn. is continued to 500° at 15° rise per min. The distillate is dissolved in pure Me_2CO in the proportion of 35 cc. per g., and the washings from the neck of the flask are added. The soln. is cooled to 0°F., this ppt. the wax, which is filtered off, redissolved in 70 g. of acetone per g. and reprecipd. at 32°F. E. E. C.

Chemical properties and examination of tar and pitches. BOCKSHAMMER. *Test u. Bitumen* 29, 20-32 (1931). —A set of tests has been arranged which allows the general characterization of unknown bituminous materials through the identification of typical components. K. H. ENGL.

Combustion and detonation in gaseous mixtures. Antidetonnants (LAFITTE) 24. The analytical distillation of gasoline (TOMREV) 2. The relative merits of gas, oil and electricity for industrial purposes (HORKINSON) 13. Critical solution temperatures of systems of SO_2 and normal paraffins (SEYER, TOON) 2. Calculating viscosity and flash point in compounded oils (KADNER) 2. Recovering oil from waste water (Swiss pat. 140,639) 14. Oil filter (U. S. pat. 1,792,854) 1. Unactivated "prefiltering medium" for treating oils (U. S. pat. 1,794,862) 1. Apparatus for extractions with liquid solvents (U. S. pat. 1,794,874) 1. Rotary impeller agitating device and filter for treating oils (U. S. pat. 1,794,916) 1. Separator for oil and water (Ger. pat. 517,425) 1. Destructive hydrodeoxygenation (Brit. pat. 338,192) 21. Wax emulsions [for manufacture of lubricating greases] (Brit. pat. 338,176) 18. Distillation of volatile substances [in refining oils] (Ger. pat. 514,740) 13. Apparatus for "neutralizing" gases from gasoline engines (U. S. pat. 1,793,813) 1.

Treating residues obtained on sweetening petroleum distillates. HERBERT G. M. FISCHER (to Standard Oil Development Co.) U. S. 1,795,278, March 3. Residue oil contg. alkyl polysulfides and obtained by the contact of sour petroleum distillates with S in the presence of a catalyst followed by distn. under reduced pressure to produce a sweetened distillate and a residue oil is heated to above 175° in order to reduce its S content. App. is described.

Removing corrosive sulfur compounds from petroleum. WARREN F. BLECKER. U. S. 1,794,668, March 3. Brine is electrolyzed in the presence of the petroleum to produce Cl and H₂, and the nascent Cl is allowed to pass directly into the petroleum while the H₂ is separately removed without introducing it into the petroleum. App. is described.

Purifying cracked petroleum vapors. CORNELIUS B. WATSON and RUDOLPH C. OSTERSTROM (to Pure Oil Co.) U. S. 1,793,885, Feb. 24. Vapors from a cracking process are cooled sufficiently to arrest conversion reactions, liquefied oil is sepd., remaining vapors are passed to a fractionating zone into the upper portion of which a material such as fuller's earth is introduced so that it falls countercurrentwise through the vapors, and vapors passing from this zone are further treated with fuller's earth or the like before passing to a sepg. zone. An arrangement of app. and various details of operation are described. Cf. C. A. 24, 489.

Apparatus (with a tank and internal pipe coil) for cracking petroleum oils. ARTHUR

H. F. W. JR. and HENRY THOMAS (to Sun Oil Co.) U. S. 1,701,200, Feb. 21. Structural details.

Cracking petroleum oils. THOMAS C. HITE (to Universal Oil Products Co.) U. S. 1,700,443, Feb. 17. In a process such as that in which the oil is heated while passing through a tube coil and thence introduced into a reaction and sep. chamber, any accumulation of unvaporized liquid oil in the reaction zone is prevented by continuously and rapidly withdrawing the liquid from the reaction zone while maintaining both the heating zone and reaction zone under superatm. pressure. Cf. C. A. 24, 410.

Cracking hydrocarbons. O. D. LEACH. Brit. 3,622,9, April 10, 1929. See Fr. 1,825,493 (C. A. 25, 1643).

Cracking hydrocarbons. PANHANDLE REFINING CO. Brit. 3,67,880, June 20, 1929. See Fr. 1,678,412 (C. A. 24, 11637).

Distilling hydrocarbons such as petroleum, tar or coal. EMIL FROXY. U. S. 1,701,542, March 3. A layer of the material is applied to a moving surface such as a heated chain or conveyor which is conducted through a distn. zone to effect distn. under heating and the distn. is positively checked by adding a further layer prior to removal of the residue from the distn. chamber (to avoid over cracking). App. is described. Cf. C. A. 24, 400.

Distilling hydrocarbon oils such as petroleum. JOHN B. THOMAS (to Sinclair Refining Co.) U. S. 1,703,670, March 3. Vapors of the oil are passed from a still to a refluxing tower and uncondensed vapors from the latter are brought into indirect heat exchange relation with fresh oil to preheat the oil and cool the vapors. The preheated fresh oil is introduced into the tower at a point below the point of introduction of vapors from the fresh oil into the tower, and additional oil is introduced into the reflux tower at a still higher point than the fresh oil vapors. App. is described. Cf. C. A. 24, 1211.

Fuel from petroleum distillation residue. RICHARD J. ROBINSON. U. S. 1,703,014, Feb. 17. Residue from petroleum distn. or cracking, in finely divided form, is mixed with a substantially neutral liquid petroleum deriv. such as residuum oil or crude petroleum and a batch of this mixt. is heated at a carbonizing temp. to form a solid, coherent fuel.

Converting hydrocarbon material into products of lower boiling point. WM. LANDIS. U. S. 1,702,912, Feb. 17. Various details of app. are described including a heating coil, expansion chamber and superheated steam supply.

Converting heavy hydrocarbon oils into products of lower boiling point. JAY J. JAKOWSKY (to C. and C. Developing Co.) U. S. 1,702,741, Feb. 17. A relatively heavy oil is vaporized under conditions such that a substantial part of the resultant vapors is in non gaseous and substantially satd. condition, and the vapors, while in operative proximity to a surface heated to a converting temp. substantially above that of vaporization, are subjected to the pptg. action of a high tension silent elec. discharge tending to drive the non-gaseous portion of the vapors toward the heated surface, and uncondensed products are conducted away for recovery of a desired product of lower b. p. App. is described. Cf. C. A. 25, 200.

Refined white viscous hydrocarbon oil. THOMAS H. ROGERS (to Standard Oil Co. of Ind.) U. S. 1,702,134, Feb. 17. A small proportion of a naphthol is added for preventing acid formation and increase of emulsifiability in use of highly refined viscous hydrocarbon oils such as those used for lubricating turbines. U. S. 1,700,135 specifies the addn. of about 0.04% of pyrogallol for the same purpose, and refers to the use generally of similar polyhydroxy aromatic compds. Cf. C. A. 24, 5149.

Separation of unsaturated constituents from hydrocarbon oils. WM. M. STRATTON (to Texas Co.) U. S. 1,702,877, Feb. 17. Material such as products from a cracking still in vapor form is continually brought into contact with a bed of adsorbent catalytic material such as fuller's earth, and adsorbed polymers are intermittently washed from the catalytic bed by the supply of solvent such as kerosene. App. is described. Cf. C. A. 24, 4028.

Apparatus for separating oil and gas. JAY P. WATKINS (40% to Guy O. Marchant and 60% to C. G. Wells). U. S. reissue 17,983, Feb. 21. Reissue of original pat. No. 1,501,683 (C. A. 22, 1470).

Modifying mineral oils. L. AYER. Brit. 337,751, April 21, 1929. Mineral oils are modified to produce viscous oils or thin gels suitable for use as lubricants by admixt. with them, while heated, of greatly modified fats, fatty oils or resin products which may be prep'd by heating in vacuo with reagents such as amines (as described in Brit. 337,742, C. A. 25, 2319) or with compds. such as benzenesulfonic acid (as described in Brit. 331,724, C. A. 24, 2944) or with compds. such as 11 sulfoxide (as described in Brit. 331,725, C. A. 24, 2901) or with saponaceous compds. (as described in Brit.

321,691 *C. A.* 24, 2910 and in Brit. 321,723, *C. A.* 24, 2911, but with a larger proportion of modifying agent than used in the previously described processes), or by use of a product formed by treating an aq. soln. of salts of amines such as benzidine or dianisidine with a soap soln. of high mol. org. acids such as the NH_4 soap of linseed oil fatty acids.

Refining mineral oils. AXTELL RESEARCH LABORATORIES, INC. Ger. 517,196, Dec. 22, 1927. See Brit. 282,738, *C. A.* 22, 3784.

Refining mineral oils. GRIGORI PETROFF. Ger. 517,197, Jan. 18, 1927. A fat or a fatty acid, e. g. stearic or oleic acid, is added to mineral oils before refining them by means of cooled or fuming H_2SO_4 . Examples are given.

Storage tank and breather system for storing mineral oils and gas. WILBUR G. LAIRD to Doberly Research Co. U. S. 1,794,263, March 3. Structural features.

Heating oil to cracking temperature and superheating steam. JOHN PAPEROS (to Foster Wheeler Corp.). U. S. 1,792,925, Feb. 17. Oil is passed serially through a portion of a heating tube bank absorbing radiant heat in a combustion chamber from a horizontally disposed tube bank in a second chamber partly sepd. from the combustion chamber at a velocity high enough at least substantially to avoid cracking of the oil being heated and heating gases are contacted with the tube system at such temp. and volume that they may be subsequently used for superheating steam to about 490° or higher. App. is described.

Oil-still operation. BENJAMIN BRODIE (to Superheater Co.). U. S. 1,794,439, March 3. Viscous material such as heavy petroleum has its viscosity reduced rapidly by heating with a heating medium such as furnace gases while the latter is at a high temp. and transfer of heat from the heating medium to the liquid is completed under counterflow conditions without permitting any vapors to sep. from the liquid during the second heating. Gaseous products are removed from the liquid after the initial heating and are added to the liquid at the end of the second heating. An app. is described comprising pipe coils in a furnace setting.

Distilling oils in vacuo. R. R. COLLINS. Brit. 337,545, Oct. 11, 1928. Distn. such as that of a heavy cylinder oil mixed with wax distillate is effected *in vacuo* by spraying in several zones across the rising vapors of the liquid in a column, collecting liquid in peripheral troughs, and recirculating in each zone by sep. pumps. Various details of app. are described.

Recovery of oils from mixtures such as destructive hydrogenation products of brown coal. MATTHIAS PIKE, RUDOLF WITTEZ and BRUNO ENGEL (to I. G. Farbenind. A.-G.). U. S. 1,794,565, March 3. Mixts. such as sludges and residues contg. oils and free C are centrifuged with an addn. of liquid aromatic hydrocarbons such as benzene possessing good solvent properties for the oil. Na isopropyl naphthalenesulfonate also may be added.

Destructive hydrogenation. J. M. JENNINGS (to Standard Oil Development Co.). Brit. 337,571, Dec. 21, 1928. In destructive hydrogenation of heavy fluid materials such as crude oils, tars, cracked residues or suspensions of coal in oils, by use of solid catalysts immune to poisoning from S, such as Cr or Mo oxides, the ratio of asphalt to catalyst in the material treated is kept below 1.5 and preferably below 0.5, in order to prevent coking. Various details of temps. and pressures and construction of app. used, etc., are given. Cl. *C. A.* 25, 589.

Destructive hydrogenation. W. R. TATE, H. P. STEPHENSON and IMPERIAL CHEMICAL INDUSTRIES, LTD. Brit. 333,150, Aug. 13, 1929. In destructive hydrogenation of carbonaceous liquids such as oils or suspensions of coal in oils, a vertical reaction vessel and a vessel for the sepn. of liquid and vaporous products are housed in the same high pressure vessel, and the separator is maintained at a temp. lower than that of the reaction vessel by using the reaction products to preheat the liquids or gaseous materials or both. The entering H may be divided into 2 streams, one of which is heated by an elec. heater and the other of which is preliminarily used to cool the reaction products and then passed to the heater. Numerous details of construction and operation of the app. are described.

Freeing oils and tars from suspended solids. I. G. FARBENIND. A.-G. Brit. 338,154, May 14, 1929. Oils and tars such as those obtained by destructive hydrogenation are freed from suspended solids by treatment with gases acid to wet litmus such as SO_2 or CO_2 and then filtering or centrifuging. A thinner such as benzine may be added and CO_2 may be bubbled through tar at a temp. of 90° .

Producing and refining cracked gasoline. FRANK A. ANGAR (to Sinclair Refining Co.). U. S. 1,795,067, March 3. Raw vapor mixt. from a cracking operation is scrubbed for sepn. of unvaporized oil, tar and tarry matter, an oil condensate of higher b. p.

than gasoline is sepl from the raw vapor mixt discharged from the scrubbing operation before the vapor mixt, including the gasoline vapors, is passed through an absorbent catalyst such as fuller's earth for rethning, and oil condensate thus sepl is then passed through the absorptive catalyst together with the vapor mixt including the gasoline vapors, to serve as a liquid washing agent and the washing agent and asocd polymers from the last mentioned step are sepl from the refined vapor mixt and are fed directly to the scrubber. An arrangement of app is described. Cf C A 24, 718.

Refining gasoline. FORDER C HERRING (to Standard Refining Co.) U S 1,705,121, March 4. Vapors of gasoline to be rethned, together with vapors of higher h. p. hydrocarbons in such proportion that the vapor mixt when condensed forms a condensate contg not less than about 50% and not more than about 75% of the gasoline product, is passed first through a condensing operation in which a part of the higher h. p. hydrocarbons is condensed, and the total vapor mixt and condensate are then passed through an absorptive catalyst and thereafter the constituents of higher h. p. than are suitable for inclusion in the final gasoline product are sepl and the remaining vapors are then condensed as the refined gasoline. An arrangement of app is described.

Filter for gasoline, etc. A. C. PENN and AMERICAN OIL CO., LTD. Brit 337,630, Sept 27, 1929. Structural features.

Filter for oil or gasoline suitable for use on motor vehicles. SIDNEY I. WOLFE (to Cuno Engineering Corp.) U S 1,702,701, Feb 17. Structural features.

Decolorizing material. WALTER S. DAVIS (to Filtrac Co. of Calif.) U S 1,702,825, Feb 17. A decolorizing clay is heated sufficiently to remove all free water and also water of crystal and then mixed with 4-8% anhyd. H_2SO_4 . The product is suitable for treating gasoline.

Method and plant for separating waxy and fatty substances from oils, particularly paraffin from mineral oil, by cooling. THE STANDARD OIL CO. Ger 517,108, July 25, 1929. This corresponds to Brit 201,146 (C A 21, 3155).

Paraffin wax. J. G. FARRINGTON A. G. (Berthold Ottens and Martin Möller-Chengul, inventors) Ger 517,240, Apr 18, 1929. To remove oil from crude paraffin wax, the latter is centrifuged while gradually warming it up to the m. p. of the pure wax.

Immersion wax. LAWRENCE W. SULLIVAN, JR. (to The Standard Oil Co.) Can 208,765, Feb 17, 1911. An immersion bath for heat treating milled articles comprises a paraffin wax and from about 0.05-0.5% of an org. aromatic compd of the class including β -naphthol, pyropallol, diphenylamine, pallic acid, hydroquinone and β -C₁₀H₇(NH₂). This additive substance inhibits the rapid formation of excessive acidity and prolongs the life of the bath.

Apparatus for determining the friction coefficient of lubricating oils, etc. VALLI S. A. P. Ger 517,211, July 2, 1929. See Brit 205,212 (C A 23, 4500).

Thickening lubricating, insulating or other oils. J. G. FARRINGTON A. G. Brit 337,634, Nov. 20, 1929. The viscosity of oils such as castor oil, whale oil or transformer oil is increased by adding esters, ethers, mixed esters or ether esters of carbohydrazes such as those of starch or cellulose, e. g., triethylcellulose, cellulose naphthenate or the like.

Purifying used lubricating oils. FURSTON, STAFFORD & Co., LTD., and J. MARRINER, Brit 317,081 2, Nov 8, 1929. Various details are described of an app suitable for use in purification processes such as described in Brit 200,771 (C A 23, 4311).

Refining used crank case oil. ROBERT A. WIGGINS U S 1,702,882, Feb 17. The oil is intimately mixed with cold concd. H_2SO_4 and then gradually heated in the course of several hrs. to about 50-60°, sludge is sepl, suspended solid matter is removed, and steam is injected into the mixt. App is described.

Emulsifying asphalt. ROBERT K. PAINTER (to Blinect, Inc.) U S 1,701,057, Feb 21. Asphalt is dissolved in a solvent such as CCl_4 and there is added to the soln an eq. amt. of colloidal clay, the solid content of which is less than 2% of the emulsion produced, the mixt is agitated (suitably at a temp. of about 55°) and alkalis of an org. acid such as HOAc are made to maintain a desired acidity, and the asphalt is added in the presence of the colloidal soln and at the proper degree of acidity. App is described.

Aqueous dispersions of water-insoluble materials such as asphalt. JACOB M. FARN (to Plinkote Co.) U S 1,701,018, Feb 21. In producing an dispersion of material such as asphalt contg less than 2% as much of a dispersing agent such as bentonite and soap or saponin, an aq. suspension is made of paste-forming colloidal material such as bentonite and the suspension is modified by adding to it an agent such as soap of such character and quantity as to reduce the surface tension of the aq. medium.

23-CELLULOSE AND PAPER

CARLETON E. CURRAN

Cellulose from sugar. V. N. IPATEV. *Compt rend acad sci U. R. S. S.* 1930A, 381.—Five g. of cellulosic material obtained from monosaccharides was analyzed. The cellulosic material proved to be sol. in Schweitzer reagent and did not reduce Fehling soln. It is seen to be cryst. in a polarization microscope, and it hydrolyzes if heated with H_2SO_4 . L. JACOBLEFF

X-ray studies of cellulose derivatives. VIII. The alkali celluloses. KURT HESS AND CARL TROGER. *Z. physik. Chem., Abt. B*, 11, 381-408 (1931).—The x-ray fiber diagram of natural cellulose disappears in fibers treated with NaOH H_2O soln. Two new fiber diagrams are obtained, one characteristic of Na cellulose I, formed in solns. contg. up to 10% NaOH by wt., and the other characteristic of Na cellulose II, formed in solns. contg. 21% or more NaOH. Drying Na cellulose II produces no change in its fiber diagram, but a new fiber diagram corresponding to an anhyd. Na cellulose III is obtained from dried fibers of Na cellulose I. Calculated fiber periods are: Na cellulose I, $10.4 = 0.2 A.U.$; Na cellulose II, $15.1 = 0.2 A.U.$; Na cellulose III, $10.2 = 0.2 A.U.$ Na cellulose I is considered responsible for the phenomenon of swelling since the greatest swelling of fibers is produced by 14.5-15% NaOH soln. Washing unstretched Na cellulose I or II with water forms a fiber giving the diagram of cellulose hydrate; by washing stretched the diagram of natural cellulose is also obtained. Washing Na cellulose I with CH_3OH produces decompn. Na cellulose II changes to Na cellulose I. Several weeks' action of NaOH CH_3OH solns. on cellulose forms Na cellulose III or Na cellulose II. If water is added, Na cellulose I or Na cellulose II are formed quickly. Sharp fiber diagrams are obtained for K cellulose (fiber period $10.2 = 0.2 A.U.$) formed by the action of 18.5% KOH soln. on cellulose, and for Li cellulose (fiber period $10.4 = 0.2 A.U.$) formed by the action of 10.9% LiOH soln. Other forms of Li and K celluloses will be described later. JANET E. AUSTIN

Fatty acid-nitrocellulose mixed esters. HANNS SCHEIDT. *Nitrocellulose* 2, 4-6 (1931).—A review of patents. E. M. STOKES

A study of some properties of nitrocellulose made from jute with special reference to its stability. RAMESH CHANDRA BAGCHI. *J. Indian Chem. Soc.* 7, 663-74 (1931).—The amt. of stabilization necessary increases with the content of impurities. Washing with dil. acid alone is not sufficient but must be followed by neutral or alk. washing. Alc. is a very effective washing agent for improving stability. Very small quantities of easily hydrolyzable pentosans and hexosans, or degraded cellulose from the action of concd. alkali, will give unstable products. The N content of the nitrocellulose decreases as impurities increase. Unless the jute is carefully purified uniform nitration is more difficult than with cotton. The Et_2O -alc. soln. of jute nitrocellulose is much less viscous than that of cotton, this indicates that the 2 celluloses are not identical. FOSTER DEB. SNELL

Tentative standard method for the determination of α -cellulose. WM. R. WILKES. *Paper Trade J.* 92, 117 (1931).—The procedure is essentially that of Method IV of the Am. Chem. Soc. (Ritter, *C. A.* 23, 5569), with the following minor modifications: (1) The use of 95% $EtOH$ or $MeOH$ followed by petr. ether or other non reacting volatile liquid for washing the α -cellulose after the final treatment with hot water is recommended to facilitate drying. (2) The detn. of residual lignin (by the Madison Forest Products Lab. method) in the α -cellulose is recommended for unbleached pulps. A. PAPINEAU-COUTURE

units of genuine beechwood lignin but that there are other alkylatable groups. The MeO groups found in the cleavage products are already present in genuine lignin. The 2 cleavage fragments identified probably originate, in part, at least, from a common unit, $-\text{OC}_6\text{H}_3(\text{OMe})_2\text{C}-$, but the MeO content of the isolated lignin and methyl lignin is not in harmony with the assumption that the lignin consists only of a chain of the units $-\text{OC}_6\text{H}_3(\text{OMe})_2\text{CH}_2\text{CH}_2\text{CH}_2-$. C. A. R.

Recent advances in the chemistry of cellulose in relation to pulp and paper. C. J. J. Fox *World's Paper Trade Rev.* 95, 374, 376, 420, 422, 462, 464, 518, 520(1931) — An address dealing with recent investigations on the chemistry and physics of cellulose, particularly as regards its structure.

Report of the committee on chemical and physical standards. H. O. KEAY *Pulp Paper Mag. Can.* 31, 185-R(1931) — The present methods of prep sheets for testing the initial strength of paper pulp prevailing in North America and in Europe are reviewed. The disintegration of the pulp, the formation of the sheet, couching, pressing, drying and conditioning are discussed. The comm. considers that the best means yet devised for the formation of the test sheet is the Brit. sheet machine, that the couching of the sheet should be done by means which eliminate the influence of the operator, which is probably best obtained by the use of a couch plate together with a roll of fixed wt. as adopted by the Brit. comm. on evaluation of paper pulp, that the work of Cameron and his co-workers (*C. A.* 21, 1370) has shown that 200 lb. per sq. in. is most satisfactory for pressing the sheets and that the procedure of the Brit. Assocn. for drying the sheets, though more elaborate, is justified because it eliminates the uncertainty which results from shrinkage when the sheet is not dried under tension. It is felt that as yet insufficient data have been accumulated to warrant definite recommendations. A. PAPINEAU COUTURE.

The determination of the strength of pulps. K. G. JONAS *World's Paper Trade Rev.* 95, 466, 468, 510-6(1931), cf. *C. A.* 25, 1074 — In order that detns. of the strength of pulp may be of value, it is essential to have a beating method and a sheet forming method, each giving reproducible results. The ideal exptl. beater should fulfil the following conditions: it must (1) work reproducibly, (2) allow of a beating effect as nearly as possible corresponding to that of the beater in practice, (3) make possible the working out of a complete beating curve without a great consumption of time, (4) permit variation of and exact regulation of the beating pressure. These requirements are claimed to be met by a lab. beater devised by J. and Kross. Six cylindrical beating containers are symmetrically arranged around the axis of a disk to which they are firmly but rotatably attached. They are closed by means of covers fitted with packing rings, each cover carrying a toothed wheel which engages with a similar wheel keyed on a fixed central shaft, so that when the machine is running each container rotates about the central shaft and about its own axis. In each container is a beating roll consisting of a free massive cylinder, in whose surface a series of half round (6 mm diam.) flutes are milled so that a tooth 2 mm wide remains between adjacent flutings. The beating pressure is regulated by adjusting the speed of rotation. The equiv. of 11.2 g. of oven-dry pulp is weighed into each beating container, the calcd. quantity of water (preferably to a consistency of 6%) is added, the material is sorted by allowing it to stand exactly 30 min. and is then disintegrated by running for 10 min. at 100 r.p.m., beating is then carried out by running at a suitable speed (150 r.p.m. or more) so as to obtain a complete beating curve within 1 hr., the containers being successively removed at stated intervals and replaced by empty containers of the same wt. A test sheet forming machine is also described, which is claimed to give very even formation and reproducible results, the variations in the strength of sheets made from the same pulp and beaten to the same degree being from 2-3 times as great with the Schopper sheet former as with the Jonas-Kross app. J. found that the highest breaking length was obtained with sheets weighing 90-2 g. per sq. m., and suggests that this wt. should therefore be adopted. A. PAPINEAU COUTURE.

Some factors affecting the strength of groundwood. A. RUBY *Papier* 33, 1283-1300(1930) — A review is given of the effects of the species of wood, conditions of growth, age, dryness and state of preservation of the wood, barking, nature, dressing and speed of the stone, grinding pressure and temp., screening and refining of the groundwood and conditions of storing.

Studies in the manufacture of groundwood pulp. H. WYATT JOHNSTON *Pulp Paper Mag. Can.* 31, 221-32(1931) — An analysis is given of the program for investigation of groundwood manufacture at the Montreal Forest Products Labs. A. P.-C.

Practice and control of heating and jordaning. F. W. MOTSCHMAN *Paper Trade J.* 92, No. 8, 127, 129(1931) — Experiences and opinions concerning the manipulation of

beater rolls and measurement of power consumption and indices of hydration. Particular attention is given to types of Jordan filter to make sp grades of paper. A P-C.

Bibliography of wood-pulp strength testing. DONALD I. CADLE. *Paper Trade J.* 92, No 9, 35-9 (1931). —This is a continuation of Moore's bibliography (C A 23, 5580) and comprises abstracts of 27 articles published in 1929 and 1930. A P-C.

Tests on cotton and wood pulp for nitrocellulose plants. M. STROUD. *Nitrocellulose* 2, 10-11 (1931). of C A 25, 413. —Methods of detn. of Cu no., degree of bleaching, neutral gum, hemicellulose, albumen, absorptive power and viscosity are given. E. M. SYMMS.

Instruments for determining the freeness or degree of beating of paper pulps. LOUIS SCHOPPER LABORATORIES. *Paper* 33, 1903-10 (1930). —The construction and method of using the latest type of Schopper Riegler freeness tester are described. A PAPERMAN COUTURE.

Hard facing in the pulp and paper industry. W. A. WISLER. *Paper Trade J.* 92, No 8, 137-13 (1931). —A discussion is given of the merits of hard facing of equipment parts subjected to wear in various applications in the pulp and paper industry. The different types of hard facing materials are indicated, together with their fields of application and the method of applying them to the parts to be faced. A P-C.

Instrument applications in pulp and paper mills cut costs and safeguard profits. REGINALD TAUTSCHOLD. *Paper Trade J.* 92, No 2, 32-8. No 4, 28-42 (1931). —A general discussion is given of the importance of instruments in pulp and paper mill operation, where they make for economical operation by permitting of a satisfactory balance and coordination of all the many steps involved. A PAPERMAN COUTURE.

The photoelectric tube in paper manufacture. S. A. STAYOR. *Pulp Paper Mag. Can.* 31, 315-6 (1931). —A brief outline is given of some of the possible uses to which the photoelectric tube can be put in the manufacture of paper. A PAPERMAN COUTURE.

Bibliography of paper making for 1930. C. J. WEST. *Paper Trade J.* 92, No 8, 145-85. No 9, 40-6. No 10, 17-54 (1931). —A classified list is given of articles on pulp and paper making published in 1930. A PAPERMAN COUTURE.

United States patents on paper making: fourth quarter, 1930. C. J. WEST. *Paper Trade J.* 92, No 10, 61-2 (1931). of C A 25, 690. —A list in numerical sequence of U. S. patents relating to pulp and paper manufacture issued during Oct., Nov. and Dec., 1930. A PAPERMAN COUTURE.

A review of current investigations at the Division of Pulp and Paper of the Forest Products Laboratories of Canada, Montreal. E. P. CAMPBELL. *Pulp Paper Mag. Can.* 31, 189-91 (1931). —A review. A PAPERMAN COUTURE.

Recent investigations of paper-making materials. ANON. *Bull. Imp. Inst.* 28, 411-24 (1930). —I. *Soft wallaba wood (Leprosia falcata) from British Guiana*. —The wood contained 55.3% cellulose on the dry basis, as compared with 39.7% in a sample analyzed some yrs. previously (C A 22, 3217). The difference in cellulose content is attributed to difference in age. Cooking with 20% NaOH on the wt. of the wood, at a concn. of 4%, for 5 hrs. at 160° yielded 46% of dry unbleached pulp and 40% of dry bleached pulp. This treatment gave a well reduced pulp, which furnished a soft, opaque, rather bulky, brown paper, of fairly good strength, but containing a no. of very small pieces of undisintegrated fiber and small black specks. The pulp bleached fairly readily to a cream color, but the black specks were not eliminated. Increasing the cooking temp. to 165° yielded 41% of dry unbleached pulp and 40% of dry bleached pulp, similar in quality to those obtained at 160°, but containing no specks. Kraft paper of fair quality, but lacking somewhat in strength, could be produced from the above pulp, but the wallaba fiber being very resistant to hydration, the cost of beating would be about twice as high as in the ordinary com. production of kraft paper. A sulfate pulping test carried out by cooking 100 g. of wood for 8 hrs. at 75 lb. and at a max. temp. of 158° (maintained for 7.4 hrs.), with 1000 cc. of liquor contg. 5.08% total and 3.48% free SO₂, yielded a gritty, resinous pulp showing very little felting power, which could not be made into sheets, this result is attributed to the high "resin" content (12% acetone sol). II. *Xerotes longifolia from Tasmania*. —The material, consisting of tough, rush like stems and rather tightly rolled thin leaves, contained 4.2% ash and 41.1% cellulose (10% moisture basis). The fibers were 1.0-3.2 mm. long (av. 2.0), and 0.005-0.020 mm. in diam. (av. 0.012). Cooking 3 hrs. at 145° with 17.2% NaOH at a concn. of 3% yielded 40% dry unbleached pulp and 36% dry bleached pulp (bleach consumption 20%), but there remained numerous particles of brown, incompletely disintegrated material, even after bleaching. Increasing the NaOH to 21.5%, the time to 3.5 hrs. and the temp. to 150° gave 39 and 35% of unbleached and bleached pulps, resp. The unbleached pulp still contained a large no. of particles of imperfectly

digested material most of which were removed on bleaching. The well beaten pulp yielded a tough opaque paper of good strength somewhat similar to that produced from esparto but less bulky, harder and somewhat stronger. III *Paupo leaves from New Zealand*. *Paupo leaves* (*Typha angustifolia* Linn var *Brounii*) contained 11.0% ash 3.0% cellulose 48.3% (43.8% on dry basis) the fibers were 0.6-3.0 mm long (av. 1.42) and 0.007-0.017 mm in diam (av. 0.0076). Cooking 7 hrs at 110° with 16% NaOH at a concn of 3% yielded 12% of dry unbleached pulp and 23% of dry bleached pulp. The test produced a fairly well digested pulp which furnished a pale brown hard rather rattly opaque paper of good strength. This showed some shrinkage on drying and contained some imperfectly sepd fibrous material which was removed by bleaching. The yield is too low to make the material promising as a source of paper pulp for the manufacture of rayon the pitch would first have to be removed from the leaves and this would still further reduce the yields. IV *Reeds* (*Phragmites vulgaris* from northern Rhodesia). The material contained 11.0% ash 1.0% cellulose 51.0% (41.1% on the dry basis) the fibers were 0.6-3.0 mm long (av. 1.7) and 0.0102-0.018 mm in diam (av. 0.01193). Cooking 4 hrs at 170° with 16% NaOH at a concn of 4% yielded 47% dry unbleached pulp and 74% dry bleached pulp, similar to that furnished by esparto yielding a fairly soft, bulky opaque paper, but requiring nearly twice as much bleaching powder as esparto. V *Sisal waste*. Cooking 5 hrs at 174° with 20% NaOH at a concn of 4% yielded 41% of dry unbleached pulp and 42% of dry bleached pulp, which gave a tough, rather rattly, opaque paper of excellent strength but containing numerous brown specks (consisting of cellular material) which were eliminated to a large extent on bleaching. A preliminary dusting treatment which removed about 30% of non fibrous matter, gave by the same treatment 53% dry unbleached pulp and 50% (equiv. to 35% on the original material) of dry bleached pulp. The unbleached pulp still contained a small amount of cellular material, which was easily eliminated on bleaching. The bleached pulp yielded a white paper of excellent strength and quality and practically free from specks. A. P. C.

Testing of paper for water resistance. PHILIP W. COWISE. *Paper Trade J* 92, No. 10 56-7 (1931).—The dry indicator method offers the greatest promise of development to a universally acceptable standard test for the degree of total size resistance. Preliminary collaborative work having given results that were inconclusive and that indicated the necessity of further development work, the following points were studied: nature and prepn. of the indicator mixt., application of the indicator on the test sheet, reading of the end point, effects of humidity and temp. A technique embodying the results of the investigation is described in detail. A. P. C.

Porosity and air space of paper. JAMES STRACHAN. *Paper Maker & Brit Paper Trade J* 81, No. 1, 21-22 (Jan., 1931).—A general discussion is given of the subject, defining the difference between porosity and air space, indicating the practical importance of the subject to paper users, outlining the methods of practical paper making adopted in controlling porosity and air space, and reviewing the methods of testing the porosity of paper. A. P. C.

Freeness testing. W. BOYD CAMPBELL. *Research Notes* 3, 25-30 (1930). *Pulp Paper Mag Can* 31, 264-7 280 (1931), cf *C & A* 24, 1976.—The principles underlying freeness testing of pulps are discussed and the conditions necessary for the operation of freeness testers are pointed out. The effect of the pressure head and the duration of flow in compacting the pulp are particularly stressed. A math analysis is given which relates the freeness as detd. by the Can freeness tester and the English sheet machine to the sp. resistance of the pulp to flow through it. For the Can tester $R = 30.44 (1000 - V)/V$, for the English sheet machine $R = 10.30 IF/W$, where R is the resistance of 1 g of pulp in a layer on 100 sq cm, V is the vol of outflow from the chamber at the end point of the test, t is the time of flow in sec., F is the fluidity of water in reciprocal centipoises, and W is the wt. of sheet formed. It is pointed out that R is different when detd. on the 2 machines and that the ratio of the 2 values is a measure of a pulp quality termed the compressibility. Typical values of the ratio for ground wood are about 1.5, while for kraft and unbleached sulfite (beaten) the value is about 3.0. It remains to be seen whether this quality varies in any one type of pulp to an extent sufficient to make the test of value for control work. A. P. C.

The folding resistance of paper. RAYMOND FOURNIER. *Papeterie* 53, 146-50 (1931).—A brief description is given of a new type of folding resistance tester, based on the same principle as the Schopper Riegler but quite different in construction. The test strip is 15 × 180 mm (same as for the Schopper tensile test) and both ends are clamped in a 2 kg wt., so that the tension is exactly 1 kg. The fold is effected by passing the test strip through a slot in a blade which is reciprocated vertically between 2

pairs of 14 mm. steel rolls mounted on ball bearings. When the strip breaks, the fall of the weight stops the motor and applies a powerful brake to the flywheel. An automatic revolutions-counter indicates the no. of double folds. **A PAPIEREAU-COUTURE**

Low-pressure steam and paper-machine drying. DONALD ROSS-ROSS. *Paper Trade J.* 92, No. 8, 106-9(1931).—The superiority of the "blow through" type of drainage system over that employing individual traps is explained and its advantages are enumerated. A blow through system of paper machine drainage, with temp. control, as installed on No. 1 machine at the Beaharmon Division of Howard Smith Paper Mills, is described in detail. It is claimed that this drying system has reduced cockling troubles, maintained more satisfactory final moisture in the sheet, and substantially increased the drying capacity. **A PAPIEREAU-COUTURE**

The efficient drying of paper. E. W. G. COOPER. *Paper Trade J.* 92, No. 8, 110-2 (1931).—A practical discussion is given of the factors involved in paper drying, including notes on the selection of felts, press rolls and the use of pre-driers. **A P. C.**

White water in paper mills. A. I. MACNAUGHTON. *World's Paper Trade Rev.* 95, 288-80, 280-8, 426, 428(1931).—General notes are given on the utilization of white water in a wide variety of mills. **A PAPIEREAU-COUTURE**

Utilizing fuel value of waste sulfite liquor. HERBERT S. KIMBALL. *Paper Trade J.* 92, No. 8, 124-7(1931).—The details of Kahles' process (C. A. 24, 600b) for the concn. and drying of sulfite waste liquor are interpreted in terms of American practice, the calens. being based on the waste liquor output of a 50-ton sulfite mill. It is estd. there can be generated 13,063 lb. of steam net per hr., at a cost of \$2.04, or about \$0.20 per 1000 lb. **A PAPIEREAU-COUTURE**

Waste liquors and gases of the paper industry. C. J. WEST. *Paper Trade J.* 92, No. 3, 48-50, No. 4, 50-6, No. 5, 46-51(1931), cf. Schroebe and W., C. A. 20, 987.—Abstracts are given of 330 U. S., Can., Brit., German Fr., Swedish Norwegian, Austrian and Swiss pats. issued during the period 1925-1930, compiled from *Paper Trade J.*, C. A. and *Papier-Fabrikant*. Author and subject indexes are appended. **A P. C.**

Color. R. CARTER. *World's Paper Trade Rev.* 95, 716-20(1931).—After a brief review of the properties, uses and advantages of basic, acid and substantive dyes and of pigment colors, the chief tests which should be applied to dyes in the paper mill are briefly discussed. **A PAPIEREAU-COUTURE**

The preparation and purification of some paper makers' pigments. E. T. ELLIS. *Paper Maker & Eng. Paper J. Annual* No. 32-4(1930).—A description is given of the prepn. and purification of bone black, ferric oxide, iron buff, lead chromate, manganese bronze, Prussian blue, red lead, smalt and ultramarine. **A PAPIEREAU-COUTURE**

The drying and coloring of paper. F. HAMILTON. *World's Paper Trade Rev.* 95, 689-94, 730-4(1931).—An address is given which deals chiefly with the properties of the various classes of coloring matters. **A PAPIEREAU-COUTURE**

Improved constant humidity room. JAMES D'A. CLARK. *Paper Ind.* 12, 1858-63 (1931).—A detailed description is given of a const. humidity room, based on, but somewhat modified from, that described by Ruff (C. A. 21, 4067), built and equipped at a cost of approx. \$1000. It has been operating with complete satisfaction since the end of 1928. **A PAPIEREAU-COUTURE**

The resins of jack pine. JOHN B. PHILLIPS. *Pulp & Paper Mag. Can.* 31, 211-9 (1931).—The resins were found to consist chiefly of resin acids and fatty and unsaponifiable matter. The fatty constituents, to which are generally attributed so-called "pitch troubles" in pulp and paper mills, are made up largely of acids and glycerides of the unsatd. type (including large proportions of elaeo and linoleic acids and traces of linolenic acid). The fats of seasoned wood were found to contain a high percentage of oxidized fatty acids, probably derived from the unsatd. acids, which oxidize and polymerize very readily. These transformations involve a change from oily, sticky material to more inert, solid products. As it is generally agreed that pitch trouble is brought about mainly by flocculation and coagulation of resin particles, it is suggested that the possibilities of tretharolamine as an agent for removing the troublesome resinous particles, or rendering them harmless, should be investigated. **A PAPIEREAU-COUTURE**

Two years' progress in corrosion resistance [in sulfate industry] (MATHESON) 9. Cr-plating of paper mill rolls (CLEVELAND) 4. Lignin, humic acid and humin (FRUCH) 10. Lignin and related compounds (HIBBERT, SANKER) 10. Destructive hydrogenation (Brit. pat. 338,192, 21. Solvents for cellulose esters and ethers (Ger. pat. 517,097) 13. Obtaining cellulose from seeds (U. S. pat. 1,794,103) 27. Synthetic rubber [for use in manufacture of threads, films, etc.] (Brit. pat. 338,182) 30. Sugars from cellu-

losic material (U S pat 1,793,081) 28 Cutting elements for paper making engines (U S pat 1,792,546) 9

Cellulose STECKBORN KUNSTSEIDF A G SWISS 142,113, May 6, 1929 Mat cellulose of good textile properties is obtained by emulsifying viscose with an agent insol therein and adding a substance sol both in this agent and in viscose. Thus, viscose is emulsified with toluene and acetone is added.

High α -cellulose product from wood cellulose ZELLSTOFFFABRIK WALDHOF and A BERNSTEIN Brit 357,445 July 29 1929 A cellulose contg 98-99% α -cellulose suitable for production of cellulose acetate nitrate or formate, etc., is obtained from bleached or unbleached wood cellulose by treatment with a 5% or stronger soln of caustic alkali or alk. carbonate in the presence of a bleaching agent such as chloride of lime. A 2.5-5.0% alkali soln gives a 90-95% α -cellulose. Various details of procedure are given.

Cooking sulfite cellulose SEACEI CHUKOVSEKH and GAIKOVA I LIACHOVETSEKH U S 1,793,264, Feb. 17. Liquor leaving the cookers is conducted to an evaporator where it is heated by fresh steam and the vapors are returned to the cookers. Various details of app. are described.

Cellulose derivatives. COURTAULDS, LTD Ger 514,892, Jan. 10, 1927. Alkali cellulose is subjected to the action of a small quantity of FeSO_4 at 25°, the free alkali is removed and the product dried and acetylated. The ethylated cellulose contains not more than one Et group to each $\text{C}_{12}\text{H}_{10}\text{O}_5$ group. Examples are given.

Cellulose derivatives. SOC ANON POUR L'IND CHIM A BALE SWISS 141,556, Aug. 31, 1928. Cellulose derivs are obtained by the action of substances of the general formula halogen- $-\text{C}_6\text{H}_4-\text{N}=(\text{C}_6\text{H}_5)_2$ on alkali cellulose. The product has basic properties which enable it to unite readily with acid dyes. Examples mention the action of chloroethyldiethylamine on cotton and cellulose. Cf C A 24, 4156.

Cellulose derivatives. SOC ANON POUR L'IND CHIM A BALE SWISS 142,173, Sept. 8, 1928. New derivs are prepd by treating cellulose with phthalic acid anhydride in the presence of tertiary bases. The product is a white mass with an affinity for cotton and basic dyes. In the example, cellulose is treated with the anhydride and pyridine at 60°.

Cellulose derivatives. SOC ANON POUR L'IND CHIM A BALE SWISS 142,175, Apr. 11, 1929. New derivs are prepd by the action of cyanuric chloride on alkali cellulose. The products are unresponsive to vat and acid dyes, but responsive to basic dyes, and can be used as the raw material for textile fibers. In an example, unbleached cellulose in KOH is treated with the chloride. Cf C A 25, 1082.

Cellulose derivatives. SOC ANON POUR L'IND CHIM A BALE SWISS 142,749, Aug. 8 1928. A new deriv is obtained by the action of phenylisocyanate on cellulose. The product can be used to make textile fibers and has a strong affinity for basic dyes.

Acetylcellulose. SAMUEL I VLES and LOUIS DE HOOF (to Algemeene Kunstzijde Unie N V) U S 1,794,126, Feb. 24. An acetylizing bath formed from Ac_2O and glacial HOAc is allowed to stand for several hrs. until part of the water in the acid has combined with the Ac_2O , and cellulose is then introduced into the bath.

Cellulose acetates, acetopropionates, acetobutyrate, etc. J G FARNENING A-G Brit 337,366 July 25, 1929. In the production of cellulose acetates, propionates, butyrates and mixed esters, the esterification is carried out in the presence of methylene or ethylene chloride as solvent instead of free acid and by effecting the reaction at temps of 20-60° degradation of the cellulose is avoided and clear solns of unusual homogeneity and viscosity are obtained, which may be spun directly or may be hydrolyzed to give acetone soly. The quantity of catalyst such as H_2SO_4 used may be reduced to about 0.5%. Various auxiliary details of procedure are described.

Precipitating cellulose acetate. K. WERNER Brit 338,214, Aug. 15, 1929. An app is described for pptn of primary or secondary cellulose acetates from HOAc soln by a pptg liquid such as water, by bringing the materials together in streams which flow into each other through a centrifugal pump to cause intimate admixt., suitably at a temp. of about 70°.

Mixed cellulose esters. C J STAUB and C S WEBBER (to Kodak, Ltd.) Brit 338,201, Feb. 18, 1929. Mixed cellulose esters, other than lacto-formate, are prepd by digesting an acyl ester of cellulose such as cellulose acetate with an aliphatic or aromatic aliphatic mono- or di basic carboxylic acid contg a hydroxy group in the α -position, such as lactic, tartaric, racemic glycolic, glyceric, mandelic and malic acids. Several examples with details of procedure are given.

Mixed cellulose esters. C J STAUB and C S WEBBER (to Kodak, Ltd.) Brit

338,202, Feb 25, 1929 Mixed esters are formed by digesting an acyl ester of cellulose such as cellulose acetate with a carboxylic acid contg a ketonic group in the α - or γ -position such as pyruvic, α ketobutyric, α ketovaleric, α ketocaproic and levulinic acids, the reaction being preferably effected in a vessel provided with a reflux condenser, in which a soln of the cellulose acyl ester and reacting acid are heated to about 100° under atm pressure. Several examples are given.

Cellulose esters and ethers. SPICERS, LTD Ger 514,945, Nov 24, 1926 Masses are formed from cellulose ethers and esters or similar carbohydrates of empirical formula $n(C_6H_7O_5)$ especially the org acid esters, by incorporating with them, hepta- or hexachloropropane alone or mixed with other solvents or softening agents Cf C A 22, 1237 2840

Solutions of cellulose esters and ethers. MAX OW ESCHINGEN (to Radio Patents Corp) U S 1,794,066, Feb 24 Materials such as nitrocellulose, cellulose acetate or ethyl cellulose are dissolved in wood oil together with other known solvents such as EtOH or MeOH (the wood oil used being the product resulting from the distn, at a max temp of 120° of residual oils obtained in the initial distn of wood)

Capsules of cellulose hydrate. RUDOLF BEYREICH Ger 514,877, April 22, 1926 The capsules are immersed in a bath of lower b p than water, e g, acetone, alc, etc, so that a complete or partial replacement of the water in the cellulose hydrate takes place by the liquid of lower b p, so diminishing the shrinkage duration of the capsule.

Capsules KALLE & Co A-G (Johus Voss, inventor) Ger 517,021, Dec 15, 1925 Addn to 515 377 (C A 25, 1672) Capsules made from viscose with which mica and (or) mosaic gold has been incorporated, are rendered opaque as described in Ger 515,377

Effecting deodorization, bleaching and sterilization of fibrous materials such as sawdust. FREDERICK W BRODERICK U S 1,792,805, Feb 17 The material is moistened with a dil NaCO₃ soln, heated to about 38°, and treated with Cl gas so that the latter reacts with the NaCO₃ present and forms HOCl directly on the material, which is washed and dried App is described

Waterproof fibrous products suitable for making milk bottles, etc. ALBERT L CLAPP U S 1,793,839, Feb 24 A thermoplastic waterproofing material such as paraffin and rosin is mixed with a suspension of wood pulp at a temp above the m p of the thermoplastic material, and while hot the suspension is made into the desired form such as milk bottles, dishes or toys

Bleaching bands, films, capsules, etc, made of regenerated cellulose or cellulose derivatives. CHEM FAB VON HEYDEN A G Ger 517,284, Dec 20, 1927 The materials are treated at atm or raised temp with solns of salts of aromatic sulfochlor amides, e g, with a 2% soln of the Na salt of *p* toluenesulfochloramide

Plastifiable material comprising acetylated wood I G FARBENIND A G Brit 337,791, Aug 7, 1929 Material such as degummed beechwood meal is treated with Ac₂O, with or without addn of concd HOAc, in the presence of 8-14% of H₂SO₄ or an equiv quantity of other acid such as HCl or H₃PO₄, at a temp below 30°, and the swollen mass which is no longer fibrous is further worked up as in cellulose acetate manuf A product may be obtained sol in HOAc or CHCl₃ Various details of procedure are given

Artificial silk. CARL HAMEL SPINN- & ZWIRNERMASCHINEN A G SWISS 142,978, Nov 13, 1929 App is described for stretching the fibers after treating them with a softening bath

Artificial silk. N V J A CARP'S GARENFABRIEKEN Brit 337,418, May 24, 1929 Various kinds of artificial silk such as that made from viscose, cellulose acetate or cuprammonium cellulose are given a dull luster by treatment with a soln of a fluosilicate, either during or after manuf, under such conditions that silicic acid is produced by hydrolysis and partly incorporated in the material so that it is not removable by water, dil acids or soap solns Various details of procedure are described

Artificial silk from viscose. N-V NEDERLANDSCHE KUNSTIJDEFABRIEK (to British Enka Artificial Silk Co, Ltd) Brit 337,350, June 8, 1929 Viscose is used with a spinning bath contg at least 30% of sulfonic acids of aromatic hydrocarbons (suitably one contg benzenesulfonic acid 37.9 and H₂SO₄ 22.3%) Glucose and alcs may be added, and the filaments may be subsequently treated with water or dil acid or salt soln, and weakly matured viscose may be used The product has an elongation of 10-13%

Artificial silk and films from viscose. I J. FRENKEL Brit 338,318, Oct 21, 1929 In order to avoid evolution of disagreeable gases during the manuf, nitrocellulose is dissolved in the viscose during its treatment with alkali (the nitrocellulose being de-

compd into cellulose and NaNO_3 being formed, which reacts with the H_2S and other noxious gases liberated in the coagulation bath) A small quantity of hydrazine is also formed and preferably converted into a harmless complex Fe compd by adding an Fe salt such as FeSO_4 . The filaments made from material thus treated are chocolate-colored when first formed but assume a normal color after desulfurization. Various details of procedure are described.

Artificial fibers from viscose. **ZELLSTOFFFABRIK WALDHOF and ARNOLD BERNSTEIN** Ger 514 712, July 14 1928. An app is described in which the coagulating and fixing liquids are sepd by a partition, the spun fiber passing through the top layers of these liquids.

Device for applying friction to filaments, etc., spun from viscose. **HERMINGHATS & Co G m b H** (Hugo Flieg, inventor) Ger 517,324, Mar 19, 1926.

Use of viscose precipitating baths containing magnesium sulfate. **JONAS J STÖCKLI and RICHARD HARTNICK** (to the firm Vereinigte Glasstoff Fabriken A-G) U S. 1,733 649 Feb 24 1928 for maintaining the compn of viscose-ppg baths contg MgSO_4 and Na_2SO_4 in soln, the soln, at a MgSO_4 content of over 220 g per l, is cooled to effect crystn of MgSO_4 in practically pure form the cryst. material is sepd, the resulting liquor is diluted and practically pure Na_2SO_4 is sepd by further lowering of the temp and the resulting mother liquor is cone'd and is added with a suitable proportion of MgSO_4 to the ppg bath to be re-used.

Artificial silk. **A Bösse** Brit 337,340, June 25, 1929. See Fr 677,479 (C. A. 24, 3113).

Artificial silk filaments from cuprammonium cellulose solutions. **BRITISH BENLERO LTD** Brit 337,608, Nov 19, 1928. In a method involving stretching and coagulation by a stream of precipitant flowing with a gradually increasing speed through a funnel within a spinning vessel, the fresh precipitant, before encountering the threads, is mixed with part of the used precipitant passing from the funnel through apertures into a surrounding space. App is described.

Liquid treatments of wet cakes of artificial silk. **J. P. BEMBERG, A-G** (to British Bemberg, Ltd) Brit 337,590, Nov 9, 1928. For treatments such as washing, or satg with soaps or fatty emulsions, etc., wet cakes of artificial silk contg acid are removed from spinning pots and placed upon supports having a yielding periphery and are then immersed in the treating baths. Various details of app and operation are described.

Apparatus for making silk filaments by the centrifugal box method. **J. G. FARBENTUNG A G** Brit 337,610, Nov 8, 1928. Impurities such as salts and acid are removed in a counter-current of washing liquor which preferably enters the guide funnel at its lower part and leaves at its upper part. Water or dil. H_2SO_4 soln may be used for the washing. Various details of app and operation are described.

Protecting rayon fibers. **ARTHUR REILLY** (to Twitchell Process Co.) U. S. 1,704 342, Feb 24 1928. Rayon fibers, after their formation and before completion of articles formed of them, are protected by coating them with inert mineral oil and "true mahogany" (a described sulfonate material).

Hollow artificial fibers. **ERSTE DEUTSCHE KUNSTSEIDENFABRIK A-G** Ger 517,355 Sept. 5 1926. Hollow artificial fibers are prep'd by incorporating into the usual spinning solns. liquid or solid substances which are insol in the solns. and which do not give rise to gases or vapors during the spinning process, the added substances being ultimately removed from the spun fibers by means of a solvent. Thus, paraffin oil may be emulsified in viscose, and after spinning washing and bleaching the oil particles may be removed from the fibers by means of C_6H_6 .

Device for the wet spinning of artificial fibers. **NOVASETA A-G ARBON** SWISS 142 046 and 142 047, Oct. 10, 1929. Addns. to 140,048 (C. A. 25, 815). Details of the funnel shaped app are given.

Apparatus for making artificial silk filaments. **COURTAULDS, LTD** Ger 517,323 July 7, 1927. See Brit 275 042 (C. A. 22, 2773).

Spinning device for artificial silk. **WEBER & Co** SWISS 142 115, Sept 10, 1929.

Spinning device for artificial silk. **CARL HAMEL A-G** Ger 514 607, Aug 6, 1929.

Spinning chamber for artificial silk. **EMIL BLASCHKE** Ger 514 844, Oct. 6, 1929.

Saccharification of wood. **EDUARD FÄRBER, GREGOIRE MINOFF and THOMAS W M POND** (to International Sugar and Alcohol Co., Ltd.) U S. 1,795,160, March 3 1929. Wood is first treated with cone'd HCl and HCl gas in quantities insufficient to effect complete conversion into sugar, then is transferred to a diffusion battery and allowed to stand without agitation, the conversion is completed by treatment with HCl of a

paper or rayon manuf., the wood is subjected to a partial vacuum to remove volatile constituents and then subjected to a relatively high gas pressure until the cells of the wood are permeated with the compressed gas, and the pressure is then suddenly released.

Apparatus (with reciprocating and rotating saws) for reducing wood to pulp. JOHN A. WIPNER (to Oswego Board Corp.) U. S. 1,795,064, March 3. Structural features.

Wood pulp. CLIFFORD F. PETERSON (to Champion Fibre Co.) U. S. 1,792,703, Feb. 17. Wood chips are digested in relatively dil. alk. cooking liquor such as NaOH and Na sulfide soln. with continuous withdrawal of liquor from the bottom of a body of liquor and chips and continuous addn. of the withdrawn liquor to the top of the same body. The temp. of the materials is raised to the desired cooking temp. by introducing steam directly into the lower portion of the body of materials, and the cooking temp. is thereafter maintained by indirect heating of the circulated liquid. App. is described.

Wood-pulp product suitable for paper making. GEORGE A. RICHTER (to Brown Co.) U. S. 1,793,194, Feb. 17. An unbeaten chemical wood pulp is homogeneously blended with like beaten or hydrated pulp. Cf. C. A. 24, 2884.

Paper pulp from groundwood. RAYMOND S. HATCH, ROBERT B. WOLF and RAYMOND P. HILL. U. S. 1,794,174 Feb. 21. Wood is ground to produce short fibered groundwood pulp. Shivers are sep'd from the pulp, the shivers are subjected to a softening and delubing treatment with a suitable soln. and the resulting pulp is mixed with the short fibered ground pulp.

Furnace for recovery of values from waste wood-pulp liquor residuum. FERNANDO FALLA (to Edge Moor Iron Co.) U. S. 1,794,006, Feb. 21. A steam boiler is operated by the furnace gases, and cooling gases are supplied to the boiler space at a point beyond the inlet of the furnace gases so that the furnace gases will be cooled to an extent sufficient to ppt. entrained "chemical ash" from them. Various details of construction are described.

"Kraft-simulating" pulp. GEORGE A. RICHTER (to Brown Co.) U. S. 1,792,510, Feb. 17. Wood chips are digested by heating under pressure in an acid sulfite liquor in which the free SO_2 and combined SO_2 are in approx. equal proportion (of 3 to 4% each), the pulp is sep'd from the digesting liquor and is further digested in an alk. cooking liquor.

Apparatus for continuous filtration and dehydration of pulp for artificial leather or paper manufacture. C. CARCANO. Brit. 337,346, July 22, 1929. Structural features.

Hydrating and beating paper pulp. G. H. CARNAHAN (to Intercontinental Rubber Co.) Brit. 338,359, Dec. 4, 1928. Pulp is hydrated by repeated squeezing and rubbing between resilient surfaces such as rubber in the presence of water (suitably by the use of rubber-covered balls in a tube mill).

Apparatus for beating paper pulp. WM. V. KNOLL (to Mid-West Machine Co.). U. S. 1,793,095, Feb. 17. Structural features.

Paper. OTTO ENGEL. Ger. 514,787, Oct. 15, 1926. Addn. to 483,735. The method of 483,735 (C. A. 24, 699), for producing layers of paper felt, cardboard, etc., with intermediate layers of fabric, gauze, bands or fibers, etc., is modified by using as the binding material, asphalt, resin, bitumen, etc., mixed with oil, tar, etc., with a m. p. considerably higher than that of the finishing material to be subsequently used.

Paper. I. G. FARBENVIND A-G. Ger. 514,856, Aug. 14, 1928. Addn. to 500,567 (C. A. 24, 4932). Cellulose material such as straw, wood, etc., is impregnated with dil. HNO_3 and treated with air after the removal of the excess of acid. The fibers are then worked up as usual.

Paper. ERNST FUES. Ger. 514,921, July 8, 1927. Paper consisting of one or more layers is impregnated, colored, glazed, etc., by sprinkling the impregnating, etc., agent on or between the paper layers in powder form in the presence or absence of diluents, treating with solid or liquid assisting agents, and pressing over rollers. Synthetic resins of the $\text{C}_2\text{H}_3\text{O}$ PhOH types are mentioned as suitable assisting agents.

Paper-making apparatus. W. P. FEENEY. Brit. 337,632, Aug. 5, 1929. Structural features.

Paper-making machine (Fourdrinier type). PAUL ERKENS. Ger. 517,229, Feb. 25, 1927. Addn. to 406,942. Structural features are described.

Metallic cloth for paper-making apparatus. RICHARD KASTNER. U. S. 1,794,624, March 3.

Filtering apparatus suitable for treating "white water" from paper manufacture. EDWIN M. BASSLER (to D J Murray Mfg Co) U S 1,793,449, Feb 17. Structural features.

Automatic apparatus for regulating the drying of paper, fabrics, etc. MASCHINENFABRIK AUGSBURG NÜRNBERG A G (Friedrich Ingert, inventor) Ger 517,332, Nov 7, 1925.

Moistening device for paper, etc., sheets. HUBERT ALTR, JOSEF DRECKS and GUSTAV RONSIEK. Ger 517,045, April 9, 1929.

Sizing paper. KARL SVEN. Ger 517,228 July 21, 1925. A thin dispersion or soln. of a sizing material, e. g., animal glue, in an amt. not exceeding 1 part of dry sizing material to 100 parts of dry paper is continuously added, at a point between the hollander and the wire cloth, to paper stuff contg. fillers, etc., and $Al_2(SO_4)_3$ or other usual precipitant. The sizing soln. may be stored for some time at atm. or reduced temp. before use.

Apparatus for locally applying coloring, etc., liquids to wet sheet material (paper). GERH. PALM, PAPIERFABRIK. Ger 517,325, Dec 21, 1924. Addn. to 509,382 (C A 25, 817).

Waterproof flexible paper. THE SCUTAN CO. Ger 514,922, Nov 25, 1927. See Brit 281,316 (C A 22, 350).

Treating paper to improve its light-reflecting properties. NATIONAL PAPER PROCESS CO. Brit 218,168, Aug 14, 1929. To cause paper or other surfaces to reflect whiter light than usual, a light filter is applied to them in uniformly distributed areas, which with the untinted areas reflects a combination of colors giving a substantially white effect (the individual tinted and untinted areas being sufficiently small to prevent conscious resolution at the normal viewing distance). The color applied (usually blue) is such as to counteract the preponderant colors normally reflected by the surface. Various details of manuf. are given.

Etching designs on paper. I G. FARBENTIND A-G (Kuno Franz and Karl Dankert, inventors) Ger 514,626, June 15, 1928. Sharp etched designs are produced by treating the outer layer of paper material with the etching dye prior to subjecting the paper to the etching process.

Variiegsted paper. LUDWIG HENNERERGER. Ger 517,137, April 7, 1929. In the manuf. of variegated paper by dropping oil colors onto soft water and then contacting the paper sheet with the surface of the water, the spreading of the color is assisted by adding a little NH_4OH to the color, in addn. to the usual ox. gall or alc. soap.

Anti-moth paper. CARTON & PAPIERFABRIK G. LAAGER. SWISS 142,372, Nov 15, 1929. An anti moth prepn. such as tar soap is added to the pulp mass during the manuf. of the paper.

Coating paper on both sides with metal foil. LEOPOLD RADÓ. Ger 514,577, April 1, 1927. Cf. C. A. 24, 4634.

24—EXPLOSIVES AND EXPLOSIONS

CHARLES E. MUNKOE AND C. G. STORM

The characteristics of an explosive: properties determined by combustion in a closed chamber. M. E. BURLOT. *Army Ordnance* 11, 283-6 (1931).—The paper describes 2 pressure gage eprouvettes, with vols. of 25 and 150 cc., for testing pulverulent explosives and smokeless powders, resp., and a special calorimeter of the Landrieu Mal-salles type, in which the eprouvettes are immersed for detg. heat of combustion. $P(t)$ is measured by registering the compression of Cu cylinders as a function of the time, by means of a pen attached to the crusher gage piston, and a revolving cylinder. For measuring gas vol. (V_g) a gasometer is connected with the eprouvette through a special needle valve. T is calcd. from measured values for P and V_g , or from the equation of combustion and the value of Q detd. by the calorimeter. Potentials of explosives are measured by this app. to an accuracy of 1/1000. C. G. STORM.

Consumption of explosives in January, 1931. W. W. ADAMS and L. S. GERRY. *Bur. Mines, Rept. of Investigations* 3093, 13 pp. (1931). ALDEN H. EMERY.

A nitro explosive five hundred years ago. F. M. FELDHAUS. *Nitrocellulose* 2, 12-3 (1931). E. M. SYMMES.

Determination of moisture in smokeless powder by the Benesch method. W. GRÄTTLICH. *Nitrocellulose* 2, 13-4 (1931).—Dry air is passed through smokeless powder.

the H_2O is absorbed, and FeOH and acetone vapors are oxidized by hot CuO , the CO_2 and H_2O therefrom being absorbed separately. F M SYMMS

A combustion theory for colloidal powders in closed vessel. HENRI MURAOUR *Compt rend* 192, 227-9(1931) —From observations on the behavior of explosive powders in a closed vessel, a theory is derived to explain the behavior based on temp., pressure and gaseous layer effects. M McMAHON

High-brisance studies IV. Penthrinit versus dynamite. ALFRED STETTBACHER *Z ges Schiess Sprengstoffw* 26, 39-40(1931) cf *C A* 25, 1998 —S answers the criticisms of Naoum (*C A* 25, 818) by describing tests in which Penthrinit is compared with various types of com blasting gelatins, the explosives varying in age from 30 hrs to 34 days. The tests were made with 250 g charges on 25 mm steel plates, 150 g charges on 7 mm Fe plates resting on an Fe support, 100 g charges on 8 mm and 10 mm Fe plates, and 30 g charges stemmed in bore holes 22.5 mm diam \times 0.4 cm deep in steel cylinders 12 cm diam \times 15.2 cm high. In every case the greater brisance of the Penthrinit mixts was demonstrated. C G STORM

Penthrinit, dynamite and the so-called high-brisance studies of Dr. Stettbacher. P NAUM *Z ges Schiess Sprengstoffw* 26, 40-5(1931), cf *C A* 25, 818 —N comments on the claims of Stettbacher (cf above abstr and *C A* 25, 1674, 1998) relative to the superiority of 'Penthrinit' mixts and their practical application. The plate tests of S are interpreted as indicating relative sensitiveness to detonation rather than brisance. The more insensitive gelatin explosives require stronger confinement than that afforded in the tests of S to develop their max brisance. Even 30 mm steel tubes do not represent the confinement of a charge in a bore hole in rock. The article is mainly controversial in nature. C G STORM

Combustion and detonation in gaseous mixtures. Antidetonnants. PAUL LAP- PITTE *Bull soc encour ind nat* 130, 15-30(1931), cf *C A* 22, 1475, 2655, 23, 2571, 24, 960, 1207. ARTHUR FLEISCHER

The testing of mine dusts. A L GODBERT AND R V WHEELER *Trans Inst Mining Eng* 80, 312-27, *Colliery Guardian* 142, 301-4(1931) —A study of some 200 mine dusts indicates that percent of ash or incombustible dust is not a true guide to inflammability. Some 16 samples contg more than 50% of incombustible matter were still classified as inflammable according to the test used. These tests emphasize the wide differences in inflammability that exist between coal dusts of different chem compn and indicate that in the use of incombustible dust as a precaution agent against coal dust, there should be greater discrimination. R S DIAN

The inflammation of coal dusts: effect of the presence of fire damp. T N MASON AND R V WHEELER. Safety in Mines Research Board, London, *Paper No 64*, 32 pp. (1931) —An account is given of inter related studies of the relative inflammability of a number of different coal dusts and the effect of the presence of fire damp in the air on the inflammability of coal dusts. There is a fairly regular relationship between the content of volatile matter in coal and the inflammability of its dust. Flame is not propagated in coal dust contg less than 12.5% of volatile matter. The presence of fire damp in the air requires the use of a higher % of stone dust mixed with the coal dust for extinguishing the explosion of the latter. When 2.5% of fire damp was present, coal dust contg only 7.5% volatile matter was just capable of propagating flame under the conditions of the expts. The tests were conducted in a steel gallery 48 in in diam and 325 ft long, in which ignition was effected by a blown out shot of 20 oz of black powder stemmed with 1 lb of coal dust from a steel cannon with 2 in bore. C G S

Gaseous explosive reactions: effect of pressure on the rate of propagation of the reaction zone and on the rate of molecular transformation. F W STEVENS *Natl Advisory Comm Aeronautics, Rept No 372*, 3-19(1930), cf *C A* 24, 1982 —A beam of light interrupted by a tuned shutter passed through a lens focused on a spark gap in the center of a soap bubble of explosive gases suspended in the center of a large sphere. Through a second window, opposite the source of light, a camera focused on the spark gap recorded the explosive transformation on a moving film. The reactions (I) $2\text{CO} + \text{O}_2 \rightarrow 2\text{CO}_2$, (II) $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$, (III) $2\text{C}_2\text{H}_2 + 5\text{O}_2 \rightarrow 4\text{CO}_2 + 2\text{H}_2\text{O}$, (IV) $1.8 \text{ Fuel} + 1.2\text{O}_2 \rightarrow 1.8\text{CO}_2 + 0.4\text{H}_2\text{O}$, (V) $1.5 \text{ Fuel} + 1.5\text{O}_2 \rightarrow 1.5\text{CO}_2 + \text{H}_2\text{O}$ were studied over a range of pressures from 100 to 2650 mm. In all reactions studied, irrespective of the order given by stoichiometric equations from 3 to 15, the rate of transformation is proportional to the pressure. This indicates a much simpler reaction than is indicated by stoichiometric equations in accord with the impact theory. The findings are strong evidence in support of the theory of chain reactions. High order reactions proceed by many simultaneous simple reactions. Previous investigators, working in const vol., have in most cases given the

rate of displacement of the reaction zone and not the rate of propagation of the zone relative to the gas it transforms. Conclusion: "The rate of propagation of the reaction zone within the gases is independent of the pressure," from which it follows that the rate of mol transformation within the reaction zone is proportional to pressure." From constant vol measurements of rate of propagation the generally noted fact that the rate of propagation is uniform for a time after ignition is due to the approach to constant conditions during this time. The work of others is discussed. A bibliography is given. A. LLOYD TAYLOR

Sensitivity of nitroglycerin: effect of static electricity on it and related compounds. H. I. NASIR, *Army Ordnance* 11, 384-5 (1931).—Sparks were passed through pieces of filter paper saturated with the liquid explosives, or through small samples on a Cu plate. The condenser furnishing the spark was of 0.003 mfd. capacity. Voltage varied from 4000 to 13,000. No explosions or ignitions were obtained in repeated tests with nitroglycerin, mixts of nitroglycerin and ethylene glycol dinitrate, propylene glycol dinitrate, nitroacetone, and 40% dynamite. With a 50% soln. of DNT in anhyd EtOH, ignition of the EtOH vapors occurred. In all tests the samples were surrounded by a cylindrical water jacket to maintain the temp. at the desired point, usually 50-60°. The insensitivity of the liquid explosives to sparks is attributed to their low vapor pressure. Conclusion: Static electricity is not a hazard on nitroglycerin lines in dynamite plants. C. G. STORM

Fire, explosion and health hazards with zapon (nitrocellulose) lacquers. MAX GRENPE, *Nitrocellulose* 2, 36-8 (1931). E. M. SYMONS

Two little-known sources of explosion. HANS WOLF, *Chem.-Ztg.* 54, 796 (1930).—Iron sulfides resulting from the corrosion of gasometers by H_2 products in illuminating gas can ignite spontaneously. Thin Al foil or Al bronze dust is capable of igniting under the influence of a strong light. A. L. HEANEY

Gas-analysis apparatus [for attachment to a miner's lamp] (Brit. pat. 338,410) 1

Poudres et explosifs. Volume mis à jour à la date du 1 janvier, 1930. Paris: Charles Lavauzelle et Cie. 199 pp.

Explosive primer composition. CHARLES H. PRITHAM, U. S. 1,704,732, March 3. Hg fulminate 37, $Ba(NO_3)_2$ 32, Sb sulfide 28, ground glass 3 and a high explosive such as trinitrotoluene 4-8 parts are used together.

Propellant explosive comprising colloidized nitrocellulose. CHARLES M. A. STINE and CHARLES E. BURKE (to E. I. du Pont de Nemours & Co.) U. S. 1,792,516, Feb. 17. Nitrocellulose is colloidized with a nitrated ester of lactic acid and a polyhydric alc. such as dinitroglycerylnitrolactate and nitroglycerylnitrolactate or glyceryldinitrolactate.

Paper shot shells. WATSON H. WOODFORD (to Remington Arms Co.) U. S. 1,793,191, March 3. A portion of the paper shell wall is impregnated with a heat absorbent comprising a lacquer plasticizer, such as various described varnish or lacquer comps. Cf. C. A. 24, 4398.

Nitrated ester of lactic acid and glycerol, etc. CHARLES M. A. STINE and CHARLES E. BURKE (to E. I. du Pont de Nemours & Co.) U. S. 1,792,515, Feb. 17. A nitrated ester of lactic acid and glycerol is prepd. by nitrating a glyceryllactate with HNO_3 and H_2SO_4 , and may be used in explosives with nitrocellulose, etc. Nitrolactates of ethylene glycol are also mentioned.

Electric blasting fuses. LAGNOZA SPOLKA AKCYJNA. Brit. 337,837, June 15, 1928. Structural features.

25—DYES AND TEXTILE CHEMISTRY

L. A. OLNEY

The constitution and properties of level-dyeing rayon dyes. KARIN SCHULZE. *Seide* 36, 43-7, 85-90 (1931). H. W. STIEGLER

Azo dyes and their intermediate products. V. Constitution and degree of dispersion of azo dyes. PAUL RUGGLI and ARTHUR ZIMMERMANN. *Helv. Chim. Acta* 14, 101-26 (1931). Cf. C. A. 25, 417.—Substantive dyes for cotton are usually colloidal in soln. In this work an attempt has been made to trace a relationship between the constitution of such dyes, their affinity for cotton, and their degree of dispersion as measured by the viscosity of a dil. soln. and the rate of diffusion through a 1% gelatin

solo Definite conclusions are difficult to form Much depends on the number and position of the SO_3H groups. A large amt of tabulated data is given which cannot well be abstracted VI. Relation between molecular size and properties of azo dyes. *Ibid* 127-41.—A series of dyes were prepd by a successive 2,6' diazo coupling of 2, 3, 4, 5 and x mols of J acid (2 amino-5-naphthol-7-sulfonic acid) All showed more or less affinity for cotton, but from a technical standpoint would be considered very poor substantive dyes J acid alone does not dye cotton The 2 J compd in soln is highly dispersed The x J compd is coarsely dispersed, nearly insol, does not diffuse, and dyes poorly The others diffuse about the same rate as Congo red and their degree of dispersion approaches that of other substantive dyes, but even the 5-J compd, with a higher mol wt (1349) than any known azo dye, can hardly be classed as a true substantive, nor a true colloid, but is on the border line L E GILSON

The resistance of coloring substances to ultra-violet light. II. GIUSEPPE A BRAVO *Boll ufficiale stir sper ind pelli mat concinnis* 8, 184-90(1930)—A continuation of earlier work (*C A* 24, 2888) Measurements are given of dyes belonging to the nitro, nitroso, azo, stilbene, pyrazolone, diphenylmethane and triphenylmethane series G. SCHWOCH

Relation between fading of dyed colors in light and their color-depth. II SOMMER *Z anorg Chem* 44, 61-5(1931)—It is well known that pale shades of dyed colors fade more strongly in light than the heavier shades By employing a Zeiss graduated photometer for the measurement of color by the Ostwald system it is possible to establish a relation between the depth of color and the amt of fading The general process to be carried out for the detn of this relation is described Graphs are given plotted on the triangulated coordinate system, and expressed in the Ostwald color units, with a dyeing of Anthracene Blue SSW Gextra on wool The relations deduced are expressed in math formulas, such as $A = n\sqrt{t}$, where A = the Ziersch degree of bleaching, t = the time of the action expressed in normal bleaching hrs, and n = the "bleaching coefficient," being a const dependent on the light fastness of the dyestuff Also the relation between the collective range of different depths of dyed color and a definitely established normal is given as $A = N\sqrt{t} (\log c - \log f)$, where N is the "bleaching constant" detd empirically from a comparison of different depths of dyed color, where c = the dyestuff quantity used and f = dyestuff quantity absorbed. The deduction of such a complicated formula from one set of expts on Anthracene Blue is assuming rather too much, but this is simple compared with further formulas involving exponential equations expressing the psycho-phys laws of Fechner with respect to color phenomena J M MATTHEWS

Chlorinating and dyeing woollen goods. J ROBERT *Boll laniera* 44, 834-5 (1930)—Treat woollen yarn for $\frac{1}{4}$ hr with 100 l H_2O , contg 15 l HCl (22° Bé), eliminate excess of bath, treat in 10-20% $\text{Ca}(\text{OCl})_2$ for $\frac{1}{4}$ hr, and again in the 1st bath for 20 min and in a cold bath with 5 g NaHSO_4 per l H_2O After washing, initiate dyeing at a lower temp than usual, and without acid in the case of easily absorbed colors, only adding it gradually until exhausting the bath For a scroop treat in the cold with 5 g soap per l, dry slightly and acidify with dil HCl , H_2SO_4 or tartaric acid For a soft grip treat 10-15 min in a cold bath with 8-10 g soap, 1-2 g olive oil, and 3-4 g ammonia (24° Bé) per l H_2O Woollen mushes absorb 120-160 g active Cl, while heavier articles take $\frac{1}{4}$ to $\frac{1}{2}$ this quantity Chlorination is often followed by a passage in Na stannate of 3-6° Bé, leaving the wool to itself for some time, and fixing with H_2SO_4 of 10° Bé, washing and drying To prevent the yellowing of whites during the steaming for fixing printing colors, the cloth is prepd with 15-20 g Na stannate per l, and dried without souring The same effect is obtained by passing in a bath with 50 g oxalic acid per l R SANSONE

The dyeing of wool with vat colors. L RINOLDI *Boll laniera* 43, 989-92(1929)—Indigo dyeings on wool sometimes have reduced fastness to light, attributed to hypo sulfite, free S, colloidal S, H_2SO_4 , or sulfites formed by decompn, to strong alkali, or acidity, to the color absorption conditions, or to the addn of glue When reduced color is absorbed, even if combined with alkali, it remains in a colloidal soln, and acts as a direct product, being absorbed mechanically as aggregated mols When absorbed as a sulfonate it acts as a well dissolved acid color, and is fixed by the wool as simple mols, much less resistant than mol aggregates Weakening of the wool is caused by NaOH, and by too much hyposulfite, which provokes, even after taking from the bath, an excessive oxidation that also destroyed a part of the color. R SANSONE

Dyeing wool with tin salt and oxalic acid. J ROBERT. *Boll laniera* 44, 837 (1930).—When cochineal or quercitron is used with aniline colors for dyeing wool

in shades faster to light and alkalis, prepare a bath contg oxalic acid 2, salt 1.5, cream of tartar 1.5 and SnCl_2 1%. Add the cochineal powder, after making into a paste with a little Sn salt soln, dye at the b p 1 hr and wash in soft water. Cu and Fe must be absent and wooden vats are preferred, when tinned Cu or Sn vats are not available. Orange II palatin scarlets, ponceaux R and the rhodamines, are used in the cochineal bath after this has been fixed on the wool, although better results are obtained by using them in a fresh bath. With the phloxins and eosins this is indispensable.

R. SANSONE

Faults in the dyeing of rayon viscose and their prevention. L. RINOLDI *Boll laniera* 44, 201-5 (1930).—Eleven shades and other faults are produced on viscose silk during dyeing by colloidal S sulfides or polysulfides coming from the manufacturing process and by ununiform threads. The S oxidizes to H_2SO_4 , weakening the fiber. Deacidification with Na_2CO_3 or NaOH is not enough for uniform shades with direct or S colors, for the fiber in the more altered parts dyes less. Better results are obtained by treating with NaOH at $70-90^\circ$ and addn to the dye bath of soap, sulfonate, pyridine, old degumming baths and glue. The greatest uniformity is obtained during the most rapid dyeing and near the b p. Colors strongly fixed in tepid baths give more regular shades at $90-95^\circ$. With dark shades, for exhausting the bath there is added, after $1/4$ hr dyeing 5-20% Na_2SO_4 . Many dyes avoid this, as it can produce uneven shades.

R. SANSONE

Dyeing unmordanted cotton in cold or tepid concentrated sea-salt baths. J. ROBERT *Boll laniera* 44, 837 (1930).—The eosins, erythrosins, phloxins, Bengal pinks and certain acid colors are dyed in a lukewarm bath contg 50 g sea salt per l. Quinoline yellow S , naphthol yellow S and metanil yellow are dyed in a lukewarm bath with 10-20 g sea salt per l. Azoflavine RR, orange II, orange GR, ponceaux R and fast ponceaux need instead a bath with 10-20 g sea salt and 0.5-2 g alum per l. The cotton is dyed for $1/4$ hr, freed from the excess of bath and dried in the cold without rinsing in water. Very bright shades are obtained with a mediocre fastness to washing. From 2 to 10 kg of color is needed for every 100 kg of cotton. The baths exhaust badly.

R. SANSONE

Conditioning textiles in laboratory experiments. W. WELTRIF *Seide* 35, 492-6 (1930).—A photographically illustrated lab app for conditioning textile samples is explained and discussed.

H. W. STIEGLER

Physical and chemical properties of textiles. I. A. J. HALL *Textile Mercury* 84, 211 (1931).—X ray methods are discussed and the micelle structures of cotton and rayons are correlated with the fiber properties of dyeing, luster and resistance to decompn by snail enzyme. The following figures give % degradation of celluloses by snail enzyme (pn const at 5.28 by Na phosphate buffer, temp 36° , time 6 days): Cotton 21, filter paper 71, cotton, solvent extracted and mercerized, 10.2, cellulose regenerated from nitrocellulose 7.8, cuprammonium cellulose 18.3, viscose cellulose 33.1, alkali sol cellulose prep'd from nitrocellulose 32.4, cellulose regenerated from acetylcellulose 70.0, alkali sol cellulose prep'd by treatment with HCl 68.5.

A. K. J.

Some physicochemical properties of the wool fiber. J. B. SPEAKMAN *J. Textile Sci.* 4, 69 (1931).—This paper relates chiefly to the adsorption of dyes by wool. Dye mols are so enormous that there would seem to be little chance of their ever entering the fiber. The size of the pores in the dry fiber were det'd by reaction with alc hydroxyl groups associated with stretching of the wool, the work required to stretch the fiber being measured first in water, and then in MeOH , EtOH , etc. The alcs employed increase progressively in mol size and wt. The resistance to extension increases until butyl alc is reached when the fiber behaves as it does in dry air. Alcs beyond butyl, therefore, do not enter the wool mol, and therefore the size of the pores in the dry wool fiber is approx. the same as the length of the n propyl alc mol. The work required to stretch wool fibers 30% (g cm per cc) in various media at 22.2° is 5.37×10^4 for octyl alc, and 1.63×10^4 for ethylene glycol. The increase of pore size of wool fibers swollen in water is more than 100%. Dry wool fibers immersed in water give an increase in length of 1% and an increase in diam of 18%. Water and other reagents are adsorbed on the surfaces of the constituent murelles, causing a swelling of the fiber. Actual detns of the pore size show the murelles are about 200 A U thick. The size of the swollen pores is about 40 A U. Thus any dye having a mol greater than this size is only adsorbed on the exterior of the fiber and will be deficient in fastness to rubbing. On transferring the dry wool fiber from dry air to water the breaking load of the fiber is changed from 22.1 to 14.9×10^4 g/cm². The increase in cross section is about 40% and the rigidity is altered in the ratio of 15:1 as compared with the breaking load. This indicates that the micelles are much longer than they are thick. The conclusion

is that the wool fiber is built up of tiny plates, 200 A U thick, the long axis of the plates being parallel with the length of the fiber. When the wool is swollen in water the distance segg the surfaces of the plates is about 40 A U, whereas in the dry fiber this is only 6 A U. J M MATTHEWS

X-ray investigations of the inner structure of wool. W T ASTBURY *J Textile Sci* 4, 1-5(1931).—The minute structural elements of the wool fiber may be studied with the aid of x rays. Wool is a complex biol structure built up from the protein keratin which is characterised by containing S in the mol. The elements of the wool fiber are to a large extent "cryst" meaning that the constituent mols are built together in a regular 3 dimensional pattern. Wool is generally called colloidal, but this is because its crystals are so minute. X ray photographs are given showing that stretched wool fibers have a different crystallographic structure from the unstretched. The latter is called α keratin and the former β keratin, though this difference is postulated entirely on the showing of the x ray photographs and is not substantiated by any chem analysis. Wool when wetted, may be stretched to twice its normal length without losing its elasticity or its ability to regain its normal length when dried. But stretched wool is very susceptible to the action of steam so that the β keratin loses its power of returning to the α form. The stretched wool is 20 times as sensitive to the action of Na_2S soln as the unstretched fiber. The x ray measurements indicate that the α to β transformation is associated with a mol elongation of from 5.15 to 10.2 A U along the axis of the fiber. Spatial mol models are given showing a theoretical structure on the wool mols. J M MATTHEWS

The oiling of wool by means of glycerol. A M SEREBRIAKOV *Niti (Suppl to Izvestiya Tekstil Prom u Torgovli)* 1, Nos 4 5, 16 7(1930). *Chimie & industrie* 25, 165(1931).—Though attempts to substitute glycerol for spinning oils have not previously been successful with high grade wools, excellent results were obtained in the present tests by using 20-30% glycerol on wool which had been incompletely scoured and still contained a relatively large amt of grease. Tests showed that Turkey-red oil also can be used satisfactorily, but its use will depend on the price. A P-C

The oxidation of the fat used in oiled wool. G SALOMONE *Boll laniera* 44, 297-300(1930).—The oils used for wool should be easily absorbed, should furnish emulsions in water with very fine fluid particles covering easily and equally the wool fibrils, should not become agglutinated in the air, should not have any corrosive action, should be very little oxidizable, should develop little heat, should be easily eliminated by a weak alkali and should give no disagreeable odor or color. The oils extd from woolen yarn stored for some time are colored brown, and have an unpleasant odor, their fatty acids turn the wool a yellow brown, not always eliminated by scouring, even with volatile solvents. Rancid oils can cause a sensible reduction in the resistance of the wool, and after scouring, a slight increase in the soly in water, in acid and in alk. solns. The more satd fatty acids are altered less easily. They give with soap solns very good emulsions, but melt too high, having a weak lubricating action. Some of their esters, e g, Me, Et or Am stearate, or Me or Et palmitate can replace natural oils and olein; they can be used alone or in mixt with oleic acid, emulsifying easily, giving emulsions with very minute particles and an increased stability. R SANSONE

Determining the size in cotton cloth. CAMILLO LEVI *Boll assoc ital chim tessile e color*, 6, 249-59(1930).—To det the amount of size, 10 g of the cotton cloth was immersed in distd water, boiled 15 min, treated in 500 cc of a 1% rapidase soln at 80-85° for 3 hrs, washed with warm distd water and cold water, boiled 15 min in 300 cc of 0.2% Na_2CO_3 soln, washed again with warm distd water and cold water, with light rubbing, washed with alc and then with ether and dried at 100-105° to a const wt. R SANSONE

Linseed oil size and viscose silk. K GÖRZE *Seide* 35, 445-9(1930).—The use of linseed oil size applied in org solvents and as an emulsion is discussed. When used in solvents the oil penetrates the fiber but as an emulsion it penetrates only slightly. Oxidation of the oil can damage the fiber if the rayon is left on bobbins for some time. It also alters the oil, making it difficult to remove, causing streaks and affinity for acid dyes, thus preventing a pure white rayon in wool rayon mixtures. Increase in wt. proceeds with oxidation. Oxidation takes place in two steps, rapid and slow. The former causes coagulation and film formation and is desired. Catalysts or ozone complete it in 24 hrs. The slow oxidation decomposes the mol into acids. Manganese and lead are used as oxidation catalysts. The lead, however, has a tendency to discolor later on when wool rayon mixts are scoured in alkali. The brown color can be obviated by adding perborate to the scouring bath. Regarding strength and elasticity, no difference was found between the soln and emulsion methods of application. Tables

Azo dyes. I G. FARBEIND A-G. (Wilfrid Hentrich and Rudolf Knoche, inventors). Ger 517,437, Mar. 31, 1929. The diazo compds from amines of the

formula $\text{RSO}_2\text{C}(\text{NH}_2)\text{CH}(\text{CH}_3)\text{C}(\text{NO}_2)\text{CH}_2$, where R is an aryl residue contg at least one sulfo group and free from or contg other substituents, are coupled with sulfonic acids of 2 aminonaphthalene or 2 amino-8 hydroxynaphthalene or their N-substitution products. Examples are given.

Azo dyes. I G. FARBEIND A-G. (Leopold Laska and Arthur Zitscher, inventors). Ger 517,438, May 8, 1929. Diazo, tetrazo or diazoazo compds not contg SO_2H or COOH groups are coupled with 2',3' hydroxynaphthoyl-1 aminodiphenyl amine or its derivs substituted in the diphenylamine residue. Coupling may be effected in substance or on a support. The products are red, brown or black dyes insol in water. Numerous examples are given.

Azo dyes. I G. FARBEIND A-G. (Hermann Chingstein and Karl Dobmaier, inventors). Ger 517,439, Dec 6, 1928. The diazo compds from o-aminobenzo-phenone and its derivs or substitution products not contg a sulfo group, are coupled with suitable components such as naphthols, arylides of 2,3 hydroxynaphthoic acid etc. Coupling may be effected in substance or on a support. The products are yellow to red dyes insol in water. Examples are given.

Azo dyes. HERMANN WAGNER and KARL BECK (to General Aniline Works). U S 1,791,210, Feb 24. Azo dyes of excellent fastness and of favorite tints are obtained by coupling, in substance or on the fiber a diazotized 2,3 aminonaphthoic acid ester or a substitution product thereof with a naphthol or with a 2,3 hydroxynaphthoic acid arylamide or a substitution product thereof. The dyes are characterized by the formula $3,2-(\text{R}_1\text{OOC})\text{C}_{10}\text{H}_7\text{N}(\text{NC}_6\text{H}_4(\text{OH})\text{R}_2)_{1,2,3}$, wherein R_1 stands for an alkyl group, R_2 stands for H or the group CONHAr and wherein the aryl nuclei may contain substituents, such as the NO_2 group or halogen, but contg no free sulfo or carboxy group. Several examples are given with details of dyeing processes. Cf C A 24, 2301.

Azo dyes. KARL HEUSNER and MAX SIMON (to General Aniline Works). U S 1,795,125, March 3. Dyes of the general formula $\text{RN}(\text{NR}_1)\text{HSO}_2\text{H}$ wherein R represents the residue of any diazo compd and R_1 a benzene nucleus, which may be substituted are obtainable, for instance, by coupling any diazo compd with a sulfamine acid of the benzene series such as the sulfamine acid of aniline, a toluidine or a cresidine. They are, when dried and pulverized, generally reddish to brownish powders dyeing wool in yellow to orange shades. An example is given of the use of acetyl p-phenylenediamine and the Na salt of 2 sulfamino-4 methyl 1 methoxybenzene as starting materials.

Azo dyes containing chromium. HUGO SCHWETZER (to General Aniline Works). U S 1,791,252, Feb 24. Products dyeing wool clear greenish yellow shades of good fastness and suitable for use as lake or pigment dyes are obtained by boiling the aq

soln of a dye of the general formula $\text{N}(\text{Cr}(\text{CH}(\text{N}(\text{NR}_1))\text{CO}(\text{N}(\text{R}_2)\text{SO}_2\text{C}_6\text{H}_4(\text{OH}))\text{CO}_2\text{H})_{2,1}$, wherein R_1 stands for an aromatic nucleus, x stands for the methyl or the carboxy group, R_2 stands for a benzene nucleus, wherein R_1 and R_2 may be substituted, and wherein the salicylic acid radical is attached to the sulfone group in o- or p position in relation to the hydroxy group of the said radical (said dyes being obtainable according to the U. S. Patent 1,685,071, C A 22, 4833) with a H_2O -sol salt of trivalent Cr, such as Cr formate, Cr acetate or Cr fluoride, for about 2-4 hrs. Examples with details of procedure are given.

Pyrazolone azo dyes. HERMANN WAGNER, HINZ EICHWEDE and ERICH FISCHER (to General Aniline Works). U S 1,794,218, Feb 24. Dyes giving yellow tints on wool are obtainable by coupling any diazo compd with a compd of the general formula

$2,4,5-(\text{HO}_2\text{S})\text{C}(\text{Me})\text{C}_6\text{H}_3\text{N}(\text{CO}(\text{CH}_2)_x\text{CX}(\text{N}))$, wherein X represents a methyl, carboxyl or a carboxylic acid ester group. These compds, which have not yet been described, are obtainable in the known manner by condensing the hydrazine with acetoacetic ester or oxalacetic ester. Several examples are given.

Producing azo dyes on acyl celluloses. FAIRBACH FELIX and WOLFGANG JAECK (to Soc anon pour l'ind chim à Bâle). U S 1,793,390, Feb 17. Products such as "acetate silk" are dyed with an azo dye of the general formula, $\text{NH}_2\text{R}'\text{N}(\text{NR}'')\text{NH}_2$, in which R' and R'' signify 2 aryl residues which may be different from each other (such as the diamino azo dye obtained by saponifying the coupling product from diazotized p-aminoacetanilide and cresidine) and the material is then treated with

nitrous acid and further coupling is effected with β -hydroxynaphthosic acid (or with β -naphthol for producing a somewhat less light fast dyeing) in a medium with a pH below 7.1. Various details of procedure are given.

Monoazo dyes. J. R. GEIGY, *Soc. Anon.* Brit. 337,577, Oct. 15, 1928. Monoazo dyes giving clear red shades, fast to light and acid or alk. lulling on wool and silk are made by combining a diazotized aryl or alkyl ether of α -aminophenol or a substitution product with a N -arylsulfo-1-amino-8-naphtholdisulfonic acid. Several examples are given.

Vat dyes. HENRICH SCHMIDT (to General Aniline Works) U. S. 1,703,138, Feb. 17, 1928. A 2-arylpyrazoleanthrone contg. in the pyrazole ring the free imino group and in the aryl residue a halogen atom in α -position to the keto group such as dichlorobenzoylpyrazoleanthrone or like compd. in which the anthraquinone as well as the benzene nucleus may contain substituents, is treated with an inorg. substance of feeble alk. reaction such as $KOAc$ or Na_2CO_3 suitable in the presence of a diluent such as $PhNO_2$ and a catalyst such as Cu powder, producing products which dye cotton in a violet vat fast blue shades. Details of procedure are given.

Vat dyes. I. G. FARBERND A. G. Brit. 337,741, July 4, 1929. Anthraquinonylaminobenzanthrones contg. 2 or more anthraquinonylamino groups, one of which is in the R ; 1 position, are converted into vat dyes by treatment with alk. condensing agents, as by fusion with alk. KOH . Numerous examples are given of the production of dyes producing brown shades in most cases.

Vat dyes. I. G. FARBERND A. G. (Fritz Baumann, inventor). Ger. 513,608, April 10, 1929. Vat dyes are produced by introducing a substituted or unsubstituted anthraquinone β -carboxylic acid residue into the amino group of α -aminoanthrimidecarbazole or its derivs. and substitution products. Thus, 5,5'-diamino-1,1'-anthrimidecarbazole (obtained by coupling 5,5'-di(benzoylamino)-1,1'-anthrimidecarbazole) is taken up in $PhNO_2$ and heated. Anthraquinone-2-carboxylic acid chloride is then added to form the dye. Other examples are given. Cf. C. A. 25, 1396, 1692.

Vat dyes. I. G. FARBERND A. G. (Walter Mieg, inventor). Ger. 514,518, July 10, 1928. Addn. to 463,835. Vat dyes of the anthraquinone series are prepd. by the action of condensing agents (especially $ClSO_3H$ or H_2SO_4) on dianthraquinonylamines which contain at least one halogen substituted acridone ring and at least one acylamino group, at low temps. The product may be oxidized. Thus, the dianthraquinonylamine deriv. prepd. as described in Ger. 513,043 (C. A. 25, 1396) by the condensation of 1 mol. of trichloroanthraquinoneacridone with 1 mol. of 1-benzoylamino-5-aminoanthraquinone, is treated with concd. H_2SO_4 . The product is treated with Na_2O_2 to give a dye which colors cotton in fast black shades. Other examples are given.

Vat dyes. I. G. FARBERND A. G. (Heinrich Nerescheimer and Georg Böhner, inventors). Ger. 517,377, Oct. 3, 1929. Vat dyes are prepd. by halogenating α -arylaminoanthraquinonealdehydes contg. at least one $ClHO$ group in α -position to an NH group, and free from or contg. other substituents. Examples are given of the prepn. of dyes from 1-phenylaminoanthraquinone-2'-aldehyde, 1-phenylaminoanthraquinone-2-aldehyde and 1,5-diphenyldiaminoanthraquinone-2'-2'-dialdehyde.

Vat dyes. I. G. FARBERND A. G. (Max A. Kunz, Karl Köberle and Erich Berthold, inventors). Ger. 517,440, July 27, 1929. Addn. to 499,169 (C. A. 24, 4168). Vat dyes contg. more than one halogen are prepd. (1) by treating dibenzanthrone or isodibenzanthrone or their derivs. with at least two different halogens or halogenating agents, simultaneously or in succession, or (2) by introducing into halodibenzanthrones or haloisodibenzanthrones a halogen other than that already present, or (3) by treating polyhalodibenzanthrones or polyhaloisodibenzanthrones so that the halogen is partly replaced by another halogen, or (4) by treating substitution products of dibenzanthrone or isodibenzanthrone so that a nitro, sulfo or other substituent group is replaced by halogen and then introducing a different halogen, if necessary. Examples are given. Cf. C. A. 25, 2003.

Vat dyes. I. G. FARBERND A. G. (Robert Berliner, inventor). Ger. 517,441, Aug. 8, 1928. Carbazole derivs. useful as vat dyes are prepd. by the action of H_2SO_4 at a low temp. on dianthraquinonylamines contg. at least two acridone rings, the products being oxidized, if necessary. An example is given.

Vat dyes. I. G. FARBERND A. G. (Felix Kaser, inventor). Ger. 517,443, April 12, 1929. Vat dyes giving yellow shades on cotton are prepd. by heating 3-mercapto-2-methylanthraquinone or its derivs. with S . An example is given.

Vat dyes. I. G. FARBERND A. G. (Karl Schumacher and Karl Eishold, inventors). Ger. 517,444, May 31, 1927. See Brit. 291,361 (C. A. 23, 1283).

Arylaminoanthraquinones suitable for dyeing cellulose esters and ethers. Soc. ANON FOUR L'IND CHIM A Bâle Brit 339,412, Dec. 22, 1928 Arylaminoanthraquinones are made by heating with H_2BO_3 and a primary aromatic amine the leuco deriv. of a hydroxyanthraquinone, or of an amino or alkylamino deriv. or of a sulfonic acid of any of these compds., in aq. soln. or suspension (or instead of the leuco derivs. the anthraquinone compds. themselves may be used in the presence of a reducing agent) If desired the reaction may be so conducted that any sulfonic groups present in the β position are completely or partially removed during the condensation. The products are acid dyes for animal fibers or for artificial fibers such as "acetate silk," or if non sulfonated, for resins, varnishes or colloids such as nitrocellulose varnish or cellulose acetate. Several examples are given of the production of dyes giving green and blue colorations.

Aniline-black dyeing on cellulose acetate. SILVER SPRINGS BLEACHING & DYEING Co., Ltd and A J HALL. Brit 337,740 July 27, 1929 Fibers, films, fabrics, etc., comprising cellulose acetate, are treated with a mixt. contg. aniline and at least one org. amine serving as an oxidation catalyst such as *p*-aminodiphenylamine, *p*-phenylenediamine, *p*-aminophenol or nitrosodimethylamine, followed by oxidation in a hot acidified soln. of a dichromate at least nearly satd. with NaCl. Other amines may be used with the aniline, such as *o*- and *p*-toluidines, *m*-phenylenediamine or α -naphthylamine. Various details of procedure are given.

Thickening agents for dyeing. FERD SICHSEL KOMM GES Ger. 514,842, Mar. 2, 1928 A thickening agent is prepared, by pulverizing vegetable mucus with a small quantity of steeped starch.

Dyes and intermediates. I G FARBERND A-G (Georg Kränlein, Heinrich Greune and Max Thiele, inventors) Ger 517,194, June 7, 1928. See Brit. 313,094 (C A. 24, 971).

Dyes and intermediates. I G FARBERND A-G Brit. 337,860 Aug. 20, 1929 Monodramic acids of 4,4'-diaminodiphenyl 1,1'-cyclohexane are made by heating with oxalic acid (preferably in the presence of a solvent) parent substances such as 4,4'-diaminodiphenyl 1,1'-cyclohexane, 4,4'-diamino-3,3'-dimethyldiphenyl 1,1'-cyclohexane, 4,4'-diamino-3,3'-dimethoxy-1,1'-cyclohexane, 4,4'-diamino-3,3'-diethoxy-1,1'-cyclohexane, 4,4'-diamino-3,3'-dichloro (or dibromo) 1,1'-cyclohexane or 4,4'-diamino-3,3'-dimethyl-4'-methylcyclohexane. The monodramic acids thus formed may be diazotized and coupled with various components to form azo dyes (of which numerous examples are given).

Dye intermediates. I G FARBERND A-G Brit. 338,240, Aug. 27, 1929 4-Aminodiphenylamine derivs. substituted in the 3- or 4' positions or both, by alkyl, Cl or alkoxy groups are obtained by condensing a *p*-alkyl, chloro- or alkoxy-substituted or an unsubstituted nitrosobenzene, with an *o*-alkyl, chloro- or alkoxy-substituted or an unsubstituted nitrosobenzene (at least one nucleus being substituted) and reducing the product (suitably with Na sulfide). Several examples are given.

Dye intermediates. I G FARBERND A-G Brit. 338,314, Oct. 18, 1929. Urea derivs. of the general formula $(NH-R-CO-NH-R'-OH)_2CO$ (in which R and R' are aromatic residues of the benzene or naphthalene series which may contain substituents but no free sulfonic or carboxylic acid group and in which at least one *o*- or *p*-position to the OH group is free) are made by condensing a diarylureadicarboxylic acid or the corresponding acid chloride in the presence of a dehydrating agent with an aminophenol. Several examples are given for producing intermediates, the alkali salts of which can be used for impregnating cotton goods for the production of azo dyes on the fiber.

Dye intermediates. I G FARBERND A-G Brit. 337,563, Oct. 10, 1929 Azo and azoxy compds. are made by reduction, in alk. soln., of compds. of the general formula $HO-aryl-NHCO-aryl-NO_2$ (in which the aryl radicals may be the same or different and may contain further substituents, and *o*- or *p*-position to the OH group being free and the NO_2 group being in *m*- or *p*-position to the CO group). Glucose, Zn dust and Na sulfide are suitable reducing agents. Several examples are given, and nitroarylacylaminophenols suitable for use as starting materials are made by condensing *m*- or *p*-nitrobenzoic acid or their halogen, alkyl or alkoxy derivs. with an aminophenol or substitution product, which has a free *o*- or *p*-position to the OH group.

Dye intermediates. I G FARBERND A-G Brit. 337,821, Aug. 8, 1929 3-Methoxytetrahydrocarbazole distd. over Pb oxide at 500-600° gives 3-hydroxycarbazole and, similarly, 2-hydroxycarbazole is obtained from 2-ethoxytetrahydrocarbazole. The starting materials may be obtained by forming hydrazones from the alkoxyphenyl

hydrazines and cyclohexanones and effecting ring closure with H_2SO_4 to the corresponding tetrahydrocarbazole

Dye intermediates. I G FARRENIN A G Brit 337,002, Sept 18, 1929 Carboxylic acids of aromatic diamino compds. in which both amino groups are nitrobenzoylated or in which one amino group is nitrobenzoylated and the other is either unsubstituted or contains a substituent which is easily eliminated or is mononitrobenzoylated aminonitro compds., are subjected to a slight reduction such as required to convert an aromatic nitro compd into an azo or azoxy compd. The resulting products may be diazotized and coupled to form azo dyes on the fiber with compds. such as 1 phenyl-3 methyl-5-pyrazolone or 8 naphthol. Several examples are given for dyeing various yellow and red shades.

Dyeing and printing fibers. I G FARRENIN A G (Irwin Hoffa and Max Kerth, inventors) Ger 514,738, Oct 21, 1927 To produce fast colors on fibers, the fibers are treated with an alkythioglycol- α -carboxylic amide of the general formula $\text{H}_2\text{NCOORSCH}_2\text{COOH}$ and an alkali. Thus, cotton is printed with a paste contg 5-chloro-1-methylbenzene-3-thioglycolic acid-2-carboxylic amide, water, NaOH, alk thickening and gum. The thickening contains gum and NaOH in the proportions 1:10. After drying, the cotton is steamed, oxidized and developed in a soln of $\text{K}_2\text{FeC}_2\text{N}_4$. The resulting colors are fast. Other examples are given. Cf C. A. 24, 733.

Dyeing cellulose esters and ethers. I G FARRENIN A G Brit 338,003, April 9, 1930 Deep greenish yellow tints are produced on products such as "acetate silk" by use of compds., non sulfonated in the nucleus, such as are obtainable by interaction of 4-amino-1,8-naphthalene anhydride with hydrazine or a deriv. or compds. of the same general type. Cf C. A. 25, 207.

Dyeing cellulose derivatives. I G FARRENIN A G Ger 514,932, Dec 19, 1929 In dyeing cellulose derivs., alkylated, arylated or aralkylated ethers or thioethers of the anthraquinone series, contg no alkyl chromophoric groups, are used. Thus, acetate silk is colored golden yellow, greenish yellow and orange by neutral soap paste contg, resp., 1,4-dimethoxy-, 1,4,5-trimethoxyanthraquinone and anthraquinone-1,4-dithiodimethyl ether.

Dyeing cellulose acetate. GEORGE H. ELLIS (to Celanese Corp. of America) U. S. 1,793,800, Feb 17 The material is treated with at least one amino compd. such as *p*-aminodiphenylamine from an aq. soln. and is then impregnated with a soln. comprising an oxidizing agent such as a mixt. contg NaClO_2 and then, without rinsing, subjected to aging in a warm moist atm.

Dyeing regenerated cellulose materials. IBERICAL CHEMICAL INDUSTRIES, LTD., and R. BRECHMAN Brit 338,111, July 3, 1929 Level dyeings are produced by use of dyes of any of the three following characters: (1) an aminobenzoylphenylenediamine of the formula $\text{NH}_2\text{C}_6\text{H}_4\text{CONHC}_6\text{H}_4\text{NH}_2$, or a homolog or substitution product of suitable character, tetrazotized and coupled with 1 mol. of a 2,8-, 1,8- or 1,3-amino-naphtholsulfonic or disulfonic acid or a *N*-substituted or azo deriv., or aminocycloaryl or aryl deriv., contg. no external amino group, of 2,3,7-amino-naphtholsulfonic acid, and with 1 mol. of a coupling component other than 2,3,7-amino-naphtholsulfonic acid or a monocyclic aryl deriv. contg. an external amino group, (2) an aminoacetanilide or a nitroaniline or a substitution product, diazotized and coupled with a component corresponding to one of those specified in "(1)" (which component, when a diazotized aminoacetanilide is used, should not contain an acylamino or unsubstituted amino group, and, when a diazotized nitroaniline is used, should not contain a nitro or unsubstituted amino group), followed by hydrolysis of the acylamino group or reduction of the nitro group, diazotizing of the aminoazo product and coupling to a component such as the other components specified in "(1)", or, (3) a diazo compd. is coupled with an amine of the benzene series which couples in *p*-position to the amino group, the aminoazo product is condensed with a nitrobenzoyl chloride, the nitro group is reduced, and the product is diazotized and coupled with a 2,8-, 1,8- or 1,3-amino-naphtholsulfonic or -disulfonic acid, or a *N*-substituted or azo deriv., or a monocyclic aryl deriv., contg. no external amino group, of 2,3,7-amino-naphtholsulfonic acid. Numerous examples are given of dyes suitable for use on "viscose silk." Cf C. A. 25, 104.

Dyeing threads, etc., with developing dyes. F. KARRER Brit 338,012, Nov. 25, 1929 The material in the form of cops or transverse wound bobbins is introduced into a vessel which is filled with grounding liquor and subjected to a vacuum after satn. with the liquid. The vacuum is broken and the liquor removed, and, without removal from the vessel, an intermediate dyeing is effected followed by development. Various details of app. are described.

Dyeing furs, hairs and feathers. KARL MARX and ERICH LEHMANN (to General Aniline Works) U S 1,795,133, March 3. In contradistinction to 4 aminodiphenylamine, the salts of the *N*-substituted-4 aminodiphenylamines are easily sol. in water, and they give good gray tints on furs, hair, feathers and the like, even when dyed from a soln. in hard water or from a soln. which is alk. as from the presence of a small quantity of NH_3 . Among the substances which may be used are: 4 amino-*N*-methyl-diphenylamine, 4 amino-*N*-ethyl-diphenylamine, 4 amino-*N*-propyl-diphenylamine, 4-amino-*N*-butyl-diphenylamine, 4 amino-*N*-benzyl-diphenylamine, 4 amino-*N*-ethoxy-diphenylamine or corresponding compds. with a substituent in the *p*-position of the benzene nucleus. The fur to be dyed may be treated with a mordant. The *N*-substituted amines may be manufd. by subjecting the *N*-substituted diphenylamines to nitration and subsequent reduction.

Dyeing and printing textiles. I G FARBENIND A-G. Brit 337,846, Aug. 13, 1929. Acid amides or esters obtainable from vat dyes or their leuco compds. as described in Brit. 324,119 (C A 24, 3380) or Brit 330,579 (C A 24, 6032) are used for dyeing or printing and development is effected with alk. agents (some of which may be added to the thickening in printing cotton so that partial development is effected by moist steaming). Several examples are given.

Dyeing and printing textiles. MANCHESTER OXIDE CO., LTD. Ger. 514,953, June 29, 1928. See Brit. 296,530 (C A 23, 2580) and Fr. 656,827 (C. A. 23, 4350).

Mordanting, weighting and metallizing textile materials, films, etc. BRITISH CELANESE LTD., W A DICKIE and F. B. HILL. Brit. 337,813, Aug. 7, 1929. Metals such as Al, Sn or Zn or alloys of lake forming and other metals, are applied (intermittently or in the form of a pattern if desired) by an app. such as a Schoop metallizing app., to textile materials, films or the like and the metal may then be oxidized on the material if not sufficiently oxidized during its application. In mordanting by this method, the dye may be applied before or after the mordant or at the same time with it, as by spraying through a stencil. In treating materials of cellulose esters or ethers, the dye is preferably used together with an org. solvent. Various details and modifications of procedure are described.

Coloring plastic masses or their solutions. I G FARBENIND A-G (Werner Möller, Karl Holzach and Hans Krzikalla inventors) Ger. 517,491, May 20, 1926. Celluloid, cellulose esters or ethers, natural or synthetic resins and other plastic masses or their solns., are colored by incorporating into them sol. Cr or Cu compds. of dyes, particularly *o*-hydroxyazo dyes. Solns. in org. solvents of the dye compd. and the celluloid, etc., may be mixed, or the materials may be kneaded together on rolls. Examples are given.

Acid baths in dyeing. I G FARBENIND A-G (Josef Nürslein and Josef Stadler, inventors) Ger. 514,801, May 16, 1926. Addn to 466,420 (C. A. 23, 529). Instead of the soap or Turkey red oil of 466,420, other soap like substances are added to the dye baths. Thus, a soln. of olein in ethylether or polyglycol is stirred with isopropylated Na naphthalenesulfonate. The product coagulates into a clear soln. and is used in acid dye baths. Butylated naphthalenesulfonic acid or its salts, or sulfite cellulose lye may be used.

Fabric printing machine. JOHN WALDRON CORP. Ger. 514,950, Aug. 9, 1928.

Dye vat and jigger apparatus. F HEWITT. Brit. 337,894, Sept. 12, 1929. Structural features.

Apparatus for dyeing tubular packages of textile threads. J BRANDWOOD. Brit. 338,134, Aug. 7, 1929. Various details of construction and operation are described.

Apparatus for dyeing thread or yarn in packages. CHARLES K. DUNLAP (to Sonoco Products Co.) U S 1,793,736, Feb. 24. Structural features.

Apparatus for passing woven fabrics through dyeing, washing or bleaching baths, etc. CHARLES G HANHAUT (to Färberei Weidmann A G) U S. 1,794,403, March 3. Structural features.

Apparatus for applying sizing, dyeing or dressing liquids etc., to, cellulose acetate or other yarns. BRITISH CELANESE, LTD., and W I TAYLOR. Brit. 337,433, May 30, 1929. Structural features.

Treating fabrics. EDUARD TSCHÖRNER and HEINRICH RIES. Ger. 514,951, Aug. 25, 1928. Pile fabric is prepd. for cutting by treating the welt with an alk. soap soln. or a soln. of alkali with sulfonated oil, to which an oxidizing agent, such as Na_2BO_3 , Na_2O_2 or H_2O_2 , is added.

Impregnating textiles. GUDULA KNÖZINGER and LEO KNÖZINGER. Ger. 517,174, Oct. 9, 1928. Textile materials are strengthened and given a weighty appearance by

applying a suspension of kaolin or BaSO_4 in glue. The treatment is particularly intended for theatrical accessories, decorations, etc.

Retting flax with yeast and sulfurous acid. MARTIN WADDELL and HENRY C. WATSON. Ger 517,103, Oct 26, 1927. See U. S. 1,708,812 (C. A. 23, 2582).

Bleaching jute fiber. J. BEALINERBLAU Brit. 337,986, Nov. 22, 1928. Jute fiber is moistened with acid such as a 2-3% HCl soln. and treated with SO_2 at normal pressure to brighten its color while retaining its natural tint and gloss.

Fibers. FIRMA A. MONFORTS Ger. 514,706, Nov. 27, 1927. Yarn, cord, etc., made from hard fibers such as manilla, etc., is smoothed and polished by treating with rubber or rubber soln. and with a mineral lubricant such as talcum.

Mercerizing fabrics C. MOCKER. Brit. 337,648, Dec. 8, 1929. The fabrics are subjected to a series of baths of increasing concns and gradually decreasing temps., detd continuously by the shrinkage or elongation of the fabric during treatment, maintained in counterflow to the movement of the fabric and suitably replenished. App. is described.

Zig-zag roller apparatus for mercerizing fabrics. O. HOFFMANN Brit. 337,712, March 10, 1930. Structural features.

Preventing creasing in liquid treatments of textile materials. BRITISH CELANESE, Ltd. and G. H. ELLIS. Brit. 338,190, Aug. 13, 1929. In treating materials such as those of cellulose acetate or other cellulose esters or ethers made by dry-spinning processes, with scouring, dyeing or other liquids, in folded form as in winch machines, creasing is avoided by first wetting out the material thoroughly in open width form. Various wetting agents are mentioned as suitable and this treatment may be combined with other treatments such as delustering with a hot soap soln.

Steaming printed textile fabrics. I. G. FARBENIND A.-G. Brit. 336,590, March 18, 1929. A steaming process as described in Brit. 333,873 (for reducing printed fabrics by passing steam from the lower part of a vessel up through the fabric surrounded with a protective filling material such as sand) is applied to various printed or dyed fabrics.

Bleaching materials such as linen. O. HOFFMANN Brit. 337,305, Feb. 9, 1929. The material is uniformly and moderately moistened and then kept in constant and uniform movement (as by passing fabric in a zig-zag path over rollers) while exposed to a low concn. of ozone (not over 2 g. per cu m. of air and preferably a much lower concn.). App. and various details of procedure are described.

Delustering artificial threads and filaments. COURTAULDS, LTD., and C. DIAMOND. Brit. 338,269, Sept. 16, 1929. Products such as those formed from cellulose acetate (suitably by dry spinning in acetone contg. 1% of Arachis oil) are treated with aq. soap soln. at or near the b. p.

Improving elasticity of knitted artificial silk fabrics, etc. COMPOLASTIC CORP. Brit. 337,400, Nov. 2, 1928. The material is stretched while moist (as after dyeing), in one direction, allowed to dry in stretched condition, and a sheet of rubber is vulcanized on one face of the fabric. The material is suitable for making waist-bands, garters, etc.

Waterproofing fabrics. BRITISH CELANESE, LTD. Brit. 338,065, Jan. 15, 1929. A material suitable for raincoats or bathing-suit bags comprises a fabric base having on it a layer of rubber compn. and also a layer of a cellulose ester or cellulose ether compn. together with a compatible synthetic resin. Various examples of materials used are given, and plasticizers, coloring agents, etc., may be added.

Waterproofing wool. MERKEL & KIENLIN GES. Brit. 338,391, Dec. 13, 1928. Waterproofing is effected by treating the material first with an aq. emulsion contg. soap and an oil or fat of animal or vegetable origin (suitably olive oil) and then with a soln. of salts of those metals the hydroxides and carbonates of which are insol. or but slightly sol. in water, such as Al formate, and subsequently heating to 60-80° for some time.

Carbonizing wool. I. G. FARBENIND A.-G. (Albrecht Kittel, Carl Daimler and Gerbard Balle, inventors) Ger. 517,201, Jan. 11, 1925. In carbonizing wool with H_2SO_4 , an aromatic sulfonic acid having a tanning action, or a salt of such acid, is added to the bath. A bath of lower H_2SO_4 content than is usual can then be used. An example is given.

Forming yarns and threads of impregnated fibers. L. S. M. LEJEUNE and J. E. C. BONGRAND. Brit. 338,381, Nov. 25, 1929. Fibers such as cotton, wool, silk, flax, hemp, ramie, etc., are impregnated with solns., suspensions or dispersions of materials such as rubber, rubber like substances or synthetic resins on the spinning-frame when the fibers are in the roving stage and before twisting into thread. Numerous details of app. and procedure are described.

Imitation linen, etc. F O MUNKTELL. Brit 338,224, Aug 20, 1929. Imitation linen, towels, handkerchiefs, etc., are made by pressing superposed layers of cellulose wadding material, which may be replaced wholly or in part by Japanese yoshino paper or the like, some of the layers of which are impregnated more strongly than others with waterproofing or sizing agents, etc., such as compounds comprising cellulose esters, gelatin, glue, paraffin oil or wax, starch, borax, Japan wax, casein, resin glue, etc., in various mixts. Gossing is usually effected simultaneously with the pressing.

Plated fabric containing cellulose acetate yarns. CAMILLE DARYEUS. U S 1,793,915, Feb 24. The material is subjected to a plating operation in the presence of a softening agent for the cellulose acetate such as an aretone soln in order to render the plating permanent.

Washing materials with soap-forming substances. A MAQUARDT and F WALTER. Brit 338,121, Aug 8, 1929. In a combined soap forming and cleansing process, the materials of soap forming character are added together or successively to the wash without previous pressing into the articles to be washed and without pressing out of the liquid and alics, hydrocarbons, colloids, ferments, bleaching agents, etc., also may be added. Carbonates may be used for saponifying fatty substances and for liberating gas which assists in the cleansing. Numerous details and modifications of treatment are described. Cf C A 24, 3969.

Edging fabrics to prevent fraying. TILLIE MILLER. U S 1,793,630, Feb 24. Various details are described of a treatment involving the application of a melted material (such as a mixt of shellac, rosin, stearic acid and Al bronze) to the edge of the fabric so that it serves as a hinder when it solidifies.

Material for "jacquard" cylinders of looms. MILES J SMITH and HERBERT I CHILWON (to Carolina Rubber Co.). U S 1,793,722, Feb 24. A compn is used formed of clean thin brown rubber 35, ZnO 10, "Dixie clay" 25, "Cyclone oil" 3, lime 8, an accelerator 0.4 and S 15 parts, cured by heating. Various details of manuf are described.

Transferring pictures to fabrics. A DENGLER. Brit 337,877, Nov 3, 1928. Pictures fast to light and to washing are applied to fabrics by a transfer comprising a paper backing carrying a paste contg colors and mordants which only commence to interact under the influence of heat and moisture applied in the transferring. K percarbonate may be added to a cooled soln of gelatin and this soln added to an aq. soln of com. Na phosphate contg indigo. Various other substances are mentioned which may be used.

Gasproof fabric suitable for aircraft gas cells. JOHN B FLOWERS. U S 1,793,075, Feb 17. Alternate layers of rubber cement and metal leaf are applied to a fabric and each layer of cement is allowed to dry until tacky before applying the metal leaf and to harden before applying the next layer of cement. The metal leaf is applied in sheets which are overlapped.

Treating fibrous materials with rubber, etc. F T. LANEY. Brit 337,359, June 10, 1929. Animal or vegetable fibers, either combed or in the form of batting, yarn or cloth are cleaned, mordanted with a latex coagulating substance, and then treated with a natural or artificial dispersion of rubber which may be assced with linseed oil, and various compounding, pore forming or other substances for making artificial leathers, fire fabrics, floor coverings, insulation, etc. Cf C A 24, 4406.

Rubber-coated fabric. JOHN R COCURE (to E I du Pont de Nemours & Co.). U S 1,795,199, March 3. A material suitable for automobile tops comprises a textile fabric coated with rubber and provided with an overlying coating of baked petroleum-asphalt varnish. Cf C A 24, 1524.

Protein products resembling wool, hair, horn, etc. SOC D'APPLICATIONS ET DE RECHERCHES SCIENTIFIQUES ET INDUSTRIELLES (S A R S I). Brit 338,015, Nov 24, 1928. Materials such as gelatin, chondrin, albumin, casein and gelose are treated with an alkali or alk. earth metal thiosulfate in the presence of an acid or salt (such as an alum) which decomposes the thiosulfate and liberates S in a colloidal or finely divided state within the mass. The material may be successively and alternately treated with reagents of this character until the entire mass is insol in hot water and swells but little, and may be subjected to a preliminary treatment or subsequent treatment with substances such as formal Na borate, a dichromate, picric or pyrogalllic acid, tannin or alum to impart insoly and impermeability to the material, and may be rendered more flexible by treatment with a final bath such as one contg glycerol, glycol, molasses or a plasticizing agent. Cf C A 24, 3868.

Protecting wool, fur, hair, etc., from "textile pests." I G FAABENIND A G. Brit 337,823, Aug 8, 1929. The materials are treated with solns of *o*-phenyl

caproylpseudothiourea, as phenylvalerylthiourea, acetylallylthiourea or other thiourea derivs of the same general type

Halogenated hydroxy-di- and tri-arylmethanes. I G FARBENIND A-G. Brit. 337,832, Aug 9, 1929. Products suitable for mothproofing wool, furs, etc., are obtained by condensation of an aldehyde deriv in which the carbonyl O is replaced by alkyloxy or inorg or org acid residues, with a *p* halogenated phenol contg a free *o*-position to a OH group but which may be substituted in the other positions by "indifferent" substituents such as halogens or alkyl groups. The products may be used as addns to dye baths. Among the starting materials which may be used are (1) methylal, methylenediacetate or the dimethylacetal of *o* chlorobenzaldehyde, condensed with *p*-chlorophenol, (2) 2,4-dichlorophenol condensed with the acetal-like chlorides and sulfates produced by adding benzal chloride, 2,6-dichlorobenzal chloride or 2-chlorobenzal chloride to H₂SO₄, (3) 2-chloro-, 4-chloro-, 2,6-dichloro-, 4-chloro-2-sulfo- or 5-chloro-2 sulfobenzal chloride condensed with *p*-chlorophenol, 2,4 dichlorophenol or mixts of *p*-halogenated phenols (4) *p*-chlorophenol condensed with the chloride obtained by treating *o* sulfobenzaldehyde with PCl₅. Condensing agents such as H₂SO₄, chlorides of Zn, Fe or Al or a mixt of HCl and glacial HOAc may be used and the condensation may be effected in solvents such as PhCl or PhNO₂. 4 Chloro- and 5-chlorosulfobenzaldehydes are prepd by treating the corresponding benzaldehyde with PCl₅.

Sulfonated hydroxyarylmethanes. I G FARBENIND A-G. Brit 337,808, Aug. 6, 1929. Hydroxy di- and tri-arylmethanes prepd as described in Brit 316,900 (C A 24, 1993), Brit 333,561 (C A 25, 607) Brit 330,893 and Brit 330,894 (C A 24, 6036), Brit 335,547 (C A 25, 1687), Brit 337,832 (preceding abstr) and Brit 338,126 (following abstr) are sulfonated (suitably in the presence of org solvents, and the products thus obtained may be used for protecting furs, wool, etc., from moths (being used conjointly with a dye bath if desired). Various details and examples are given.

Brominated and chlorinated *o*-hydroxy-di- and tri-arylmethanes. I G FARBENIND A-G. Brit 338,126, Aug 9, 1929. Comps of this character are prepd by brominating or chlorinating the corresponding parent materials, and are suitable for use in mothproofing wool, fur, hair etc., and for this purpose may be added to dye baths. Examples are given of the prepn of 2,2'-dihydroxy-5,5'-dichlorodiphenylmethane, 2-hydroxy 3,5-dibromodiphenylmethane, 3,5-dichloro-2-hydroxydiphenylmethane and some similar comps.

Wetting agents. I G FARBENIND A-G. Brit 337,737, April 27, 1929. Monoamines contg only one hydroxyalkyl group and free from aromatic groups are condensed with aliphatic carboxylic, sulfonic or sulfonated carboxylic acids contg at least 8 C atoms or their esters or halides (at a temp above 100° when the free acids or their esters are used) and the products may be treated with alkylating agents and (being usually basic in character) may be converted into salts with acids. Numerous details and examples are given for prepg comps which may be used as wetting, emulsifying, cleansing and dispersing agents in the textile industry, etc. Cf C A. 24, 1947, 4906.

Apparatus for degreasing clothes by treatment with solvents. MAURICE DEW HIRST (one half to Leroy A Goodwin). U S 1,795,170, March 3. Structural features.

Bag filter for filtering garment-cleaning solvents or other liquids. ANTON L DORFNER. U S 1,794,281, Feb 24.

Apparatus for purifying garment-cleaning liquids by materials such as saponifying agents and water. CLARENCE V FUQUA (to Cleaners Equipment Corp.). U S 1,793,475, Feb 24. Structural features.

26—PAINTS, VARNISHES AND RESINS

A. H. SABIN

Additional field tests on quick-drying paints. H A GARNER AND L P HART. *Am Paint & Varnish Mfrs' Assoc Circ* No 377, 113-9(1931).—The formulas of 20 more paints supplementing those reported in C A 25, 1104 are given. G G S.

Rapid methods for detecting with organic reagents the metals used in the paint industry. HANS SCHEIMER. *Farbe u Lack* 1931, 111-2.—Methods are compiled for the detection of Fe, Al, Mn, Co, Zn and Mg. G G SWARD.

Experiences in painting façades (plaster). HANS WOLFF. *Farbe u Lack*, 1931, 99-100.—Regardless of the type of material (aq or non aq) used to paint plaster, water must be excluded after the coating is in place. If water does not destroy the film

itself, it promotes efflorescence, permits growth of molds, collection of soot, etc. Spraying a painted surface with a colloidal soln of paraffin (Schmidt's encaustic process) is effective in excluding water. G G SWARD

The preparation of precipitated white lead with high hiding power. A I KOGAN *Farben Ztg* 36, 826-8(1931)—White lead prep'd by pptn from a basic Pb acetate soln by CO_2 is low in hiding power because of the formation of considerable PbCO_3 . When CO_2 is passed through a basic lead acetate soln, $21\text{Pb}(\text{OH})_2 \cdot \text{PbCO}_3$ is formed in the early stages. Later PbCO_3 is also formed at the expense of some of the basic carbonate already formed. If the process is stopped in the early stages a pigment of high hiding power is obtained. This ppt may be removed and the process repeated several times. Hiding-power values for the pigment obtained in the improved manner ranged from 22.4 to 29.3 sq cm/g as compared with 15.3 sq cm/g for the complete pptn method. The sp gr ranged from 5.95 to 6.26. G G SWARD

The red lead question. A V BLOM *Farben-Ztg* 36, 1015-6(1931)—The reaction of linseed oil with red lead is governed not only by the particle size and percentage of PbO , but also by the cryst. structure of the PbO . E. g., flat plates and spheres are less reactive than irregular crystals or clusters. The form which is ideal for non reactivity is a kernel of PbO surrounded by small spheres of Pb_2O_3 . G G SWARD

Adulteration of red lead and chrome green. F. DINSLAGER AND B. STENPEL, *Pharm Ztg* 76, 262(1931)—Red lead (minium) was found admixed with heavy spar (BaSO_4) colored with an aniline dye. Chrome green (normally a mixt of oxides of Cr) carried PbCrO_4 and BaSO_4 . W. O. E.

The colorimetry of pigments and a suggested scale of fastness. G. F. NEW, G. S. DISNEY AND D. L. TILBEARD *J Oil & Colour Chem Assoc* 14, 3-48(1931)—A method for quantitatively measuring the extent of fading of colored surfaces is described. The color quality before and after fading is located on the trichromatic color chart. The length in color triangle units (C. T. U.) of the line connecting these loci is then plotted against the time of fading. In most comparisons brightness need not be considered. Colors of different hues, satn. and brightness may be directly compared. According to their fastness pigments and dyestuffs may be placed in 10 or more groups, group 1 being the most stable with a fading line of 1 to 1.33 C. T. U. for 75 hrs' exposure to a C arc. The lengths of the fading lines of successive groups bear a ratio of 3:4. Examples of the fading of red pigments diluted with white pigments are given. The study of the color characteristics of pigments is best made at a dilution of 9 parts of diluent. The medium recommended is a 1% soln of gelatin in H_2O in such quantity that not more than 0.5% of the film is gelatin. A discussion follows. G G S

Fastness of pigment finish to rubbing. M. C. LAUB AND R. DENVER, *J Intern. Soc. Leather Trades Chem* 15, 107-13(1931)—The whole question of the manifold pigment finishes is critically considered. Suitable grinding media for pigments are soap solns., sulfonated oils, and solubilized mineral oils, pigments ground in soap soln possess exceptional covering power and stability in suspension. Pigments should not be allowed to dry before mixing with the binder. NH_4OH is preferred to horax or soda for dissolving casein or shellac to be used as a binder, since the excess NH_3 is evaporated on drying and the resulting deposit is more waterproof. Since a finish contg. only shellac-casein ammonia pigment would be too brittle, various gums are added. Casein (and other protein binders) are made insol. by HClO , but the resulting finish is brittle, and available softening agents are not very satisfactory. Tung oil seems to be an excellent binder. The following tentative procedure is suggested: Add 100 parts tung oil to 10 parts soap dissolved in 250 parts water (not over 60°), emulsify, let stand 24 hrs., add 250 parts H_2O , emulsify, add 100 parts dry pigment and grind until the pigment is sufficiently fine. Mix this pigment emulsion with (1) a soln contg. 15% casein and 1% conc. NH_4OH , (2) a 10% soln of glue and (3) glycerol, in the proportions 15:30:25:8. Sterilize with about 0.13% nitrobenzene. Apply by spraying, and top off with HClO soln. H B MERRILL

Pigments and lacquers from organic coloring materials. A. SANDER *Industria chimica* 5, 1490-5(1930)—A brief summary of the history, classification, theory, prepa., use and analysis of org. dyestuffs. A. W. CONTIERI

Reactions between inorganic basic pigments and linseed oil. K. CHARISIUS AND E. KRINGSCHER *Farben Ztg* 36, 780-3(1931), cf *C. A.* 24, 3913-4—Equal apparent vols. of ZnO , basic Pb carbonate, basic Pb sulfate, highly disperse red lead, ordinary red lead and litharge were each suspended in raw linseed oil in the absence of air. The reaction of the oil with the pigments was essentially the same as in systems in which air was not excluded, viz., large amts. of metal were found in the oil in ordinary red lead and litharge and relatively small quantities in the other pigments. G G S

Some properties of pigment-oil pastes. W. DROSTE. *Farben-Ztg.* 36, 916-8 (1931), cf. C. A. 25, 831.—The relationship between the oil absorption of pigments and the consistency of paint is reviewed.

Application of a few colloid-chemical methods of investigation to the drying of linseed oil. H. FREUNDLICH AND H. W. ALBU. *Z. anorg. Chem.* 44, 56-60 (1931).—Depolarization expts. as well as viscosity measurements and ultramicroscopic examn. show that at a low drying rate colloidal properties are not obtained in pure or saccharine linseed oil.

The drying time of Linseed oil. J. RINSE. *Chem. Weekblad* 28, 131-2 (1931).—The influence of the humidity and other factors of the atm. on the drying time of linseed oil is discussed.

Tung-oil tree. WILSON NEWELL, HAROLD MOWRY AND R. M. BARNETT. Fla. Agr. Expt. Sta., Bull. 221, 5-63 (1930), cf. C. A. 24, 1526.—The Fla. acreage in tung-oil trees is about 5000. The fruits on examn. gave 55% hulls, 45% meat, 49% oil in the meat and 32-9% of oil in the seeds. Fully matured and thoroughly air-dried fruits consist of approx. 56% seed and 44% hulls. The tree thrives best in an acid, well-drained soil. Lack of N in the soil is the principal limiting growth factor in Fla. but a well-balanced mixed fertilizer gives max. growth and yields of nuts. A tung-oil expression plant is located at Gainesville. The oil is graded as cold Pressed Fla. No. 1 and Fla. No. 2. The appearance of the former is pale and clear with a sp. gr. at 15.5° of 0.941, acid no. 0.5, sapon. no. 193.8, I no. 160, A. S. T. M. heat test 10.75, and η_{sp} 1.519. Hulls and pomace remain after extrn. of the oil. Pomace contains about 6% of N and compares favorably with castor-bean pomace as a fertilizer. The hulls contain a little N and K and are valuable chiefly as a mulching material. For varnish making, the oil gives a clearer and faster drying finish than the ordinary imported Chinese oil.

Reactivity of mineral umbers with oils. A. EIBNER, RICHARD SCHWAB AND E. ROSSMANN. *Farben-Ztg.* 36, 900-1, 1006-7, 1046-9 (1931).—The formation of soaps in mixts. of pigments and drying oils is usually due to the presence of the original fatty acids. Only a few pigments, e. g., red lead, are able to react with glycerides under the conditions of storage and drying usually prevailing. The principal requirement is an elevated temp. If a reaction does occur, the products are mixts. of normal soaps and mixed glycerides of the metal, the so-called aggregates. Such a reaction should not be called sapon. With 2 umbers, no reaction took place with either linseed oil or poppyseed oil until resins were added. This indicates greater reactivity of resin acids than of fatty acids. Metal aggregates improve not only the drying but also the durability of paints. Under certain conditions linseed and poppyseed oils, and presumably all other drying oils, yield H₂O-soluble cryst. compds. of a peroxide nature. With further oxidation, the cryst. character is destroyed. Such a reaction in the drying of oils is contrary to Auer's theory.

Lacquer poisoning. G. OUDENORGE. *Deut. woch. Wochschr.* 56, 961-3 (1930).—Several cases of poisoning among workers in a hat factory were traced to the action of *tetrachloroethane*.

Nitrocellulose lacquers in the wood-working industry. PHILLWITZ. *Nitrocellulose* 2, 8-10 (1931).

Alcohol-soluble nitrocellulose. HANS WOLFF AND B. ROSEN. *Farben-Ztg.* 36, 964-6 (1931).—The hardness, abrasion resistance, tensile strength, elongation and resistance to pressure of a no. of alc.-sol. nitrocelluloses were detd. Better properties were exhibited by those samples possessing the highest viscosities.

New method for examination of pictures. A. M. DE WILD. *Chem. Weekblad* 28, 121-8 (1931), cf. C. A. 24, 1739.—A summary of the scientific methods of picture examn. developed during the last 3 years. Chem. analysis can det. with certainty the period in which a painting was made. Röntgen-ray examn. is a valuable method for the detection of technical particulars about a certain painter. The method of magnified photographs and photography with grazing light need too much of a subjective interpretation yet to give conclusive and convincing results.

Determining the roughness of a surface. E. WERTH. *Farben-Ztg.* 36, 875-6 (1931).—Since the amt. of paint required to cover a surface depends upon the roughness as well as the porosity, a method of detg. roughness is desirable. W. does this by making an impression of the surface on timoil.

Rosin used to produce rosin oil. V. E. GROTLISCH. *Proc. Am. Soc. Testing Materials* 30, Pt. II, S34-S (1930).—Rosin oil is produced by distg. 6500-12,000 lb. of rosin, usually of a dark or opaque grade, in a cast Fe or steel still. Water is first liberated. Rosin spirits (pinolin) begin to distil at 160°, "light oil" at 200°, "kidney

oil' at 250° and 'bloom oil' at 270°. The "kidney oil" is the most valuable fraction. It is used in the manuf. of greases and printing inks and may be redistd. to give the higher grade oils. Other uses of rosin oils are listed. "Synthetic" rosin oils are sold of rosin in mineral oils. H K SALZBERG

Rosin in the manufacture of rosin esters. STEPHEN BARCOCK. *Proc Am Soc Testing Materials* 30, Pt II, 705-7(1930).—The discussion is limited to ester gum, its history and manuf., and the properties of rosin which are important in its manuf. Wood oil-ester gum combinations are popular because they require relatively little time to age. Clean rosin of a clear grade having a high m p and acid no. is desirable in manuf. ester gum. H K SALZBERG

Tests on oil cloth and artificial leather. R HESS. *Nitrocellulose* 2, 6, 8(1931).—Detailed methods of proximate analysis are given. I M SIMMONS

Rosin as a linoleum component. M K BAAR. *Proc Am Soc Testing Materials* 30, Pt II, 803-6(1930).—Since 1926 the use of rosin by the linoleum industry has increased by 31%. At present this industry ranks fourth in the consumption of rosin. Linoleum cement, which is made from wood rosin, accounts for 90% of the total rosin used by the industry. The remainder gum rosin goes into linoleum paint. Linoleum binder consists of oxidized linseed oil rosin and other resins. The rosin serves as a solvent for the oxidized oil, its solvent power being greater the higher its acid no. A rosin of very high acid no. cannot be used however, because of its high m p and susceptibility to attack by alkali. The presence of rosin in the binder gives it heat sensibility enabling it to be worked under the compounding machines. Rosin also serves to keep the finished linoleum in a flexible condition. Substitutes for rosin are deficient in this stabilizing effect. H K SALZBERG

Fire, explosion and health hazards with nitrocellulose lacquers (GUTHRIE) 24. The resins of jack pine (PHILLIPS) 23. Solvents (DURRANS) 2. Solvents for resins (Ger pat 517,097) 13. Arylaminoanthraquinones [dyes for resins and varnishes] (Brit pat 338,412) 25. Coloring plastic masses or their solutions (Ger pat 517,491) 25. Synthetic rubber [for addition to paints or lacquers] (Brit pat 338,152) 30. Discharge device [apparatus for atomizing paints with CO₂] (U S pat 1,704,185) 1.

BIANCHI, CALISTO. *Celluloseesterlacke*. Revised German ed by Adolf Weihe. Berlin J Springer 328 pp bound, M 22.50. Reviewed in *Chem Trade J*, 88, 261 (1931).

Paints. A ROEHL. Brit 337,523, Sept 5, 1929. Paints which dry rapidly are made by heating, drying or semi-drying oils such as linseed, poppyseed or wood oils, with metal oxides such as those of Zn or Ti, to about 100-150° until a pasty product is formed and mixing the latter with solvent and pigments. Resins may be added.

Paints containing finely divided metals. L V ADAMS (to British Thomson Houston Co., Ltd.) Brit 337,682, Jan 11, 1929. Metal powders such as Al, Zn, Cu alloys (bronzes), Ag or Au are suspended in a soln comprising a volatile solvent and a synthetic resin compn. such as one formed from glycerol, phthalic anhydride, linseed oil fatty acids, wood oil resin and diethylene glycol. Paints thus formed contg Zn or Al may be applied to base metals such as iron or Cu and the Zn or Al subsequently alloyed with the coated base metal surface by heating.

Apparatus for mixing ingredients of paints, etc. I I BREGAN. Brit 337,945 Oct 21, 1929. The materials to be mixed are passed under pressure through a pipe contg a roll of metal net (suitably of 25-35 mesh and about a meter long).

Pigment. SOC ANON ROUX L'IND CHIM A BALE. Swiss 141,886, Aug 21, 1928. A colored pigment is prepd by mixing chalk or other similar Ca compds with the azo dye produced by the action of COCl₂ on the vapord dye from the diazotized p-toluenesulfonic acid ester of 1,8-ammonaphthol-3,6 disulfonic acid and cresidine, 1,4,3-Me (MeO)(NH₂)C₆H₃. The resulting pigment is red and has good light, water- and weather-resisting power. Examples are given.

Iron oxide pigment. JULIUS LAUX (to I G Farbenind A G). U S 1,703,041, Feb 24. An aromatic nitro compd such as PhNO₂ is reduced with Fe and an aq soln contg at least 25% of a salt of an org amino compd such as PhNH₂ HCl, and the resulting Fe oxide sludge is sepd from the aromatic amino compd produced and is purified by washing. U S 1,703,942 relates to a process in which an aromatic nitro compd such as PhNO₂ is reduced with Fe and an aq soln of acid such as HCl contg at least 2 mols of hyd acid in 1000 parts of soln, insufficient in quantity to dissolve the Fe,

the sludge formed being sepd from the resulting amino compd. and purified by washing. Cf. C. A. 24, 5171

Drying oil. IRVIN W HUMPHREY (to Hercules Powder Co) U. S. 1,793,220, Feb. 17, 1929. A drying oil suitable for use instead of linseed oil comprises a polymerized terpene product composed mainly of polymers having a higher h. p. range than that of the polymer ($C_{15}H_{22}$).

Modifying the physical properties of fatty oils. L. AUER Brit. 337,736, April 24, 1929. For producing a modified product sol. in a desired org. solvent for the manuf. of "leather cloth," lacquer or varnish, the fatty oil is treated with a modifying agent insol. in the selected solvent, and there is heated with the material over 2% of a soly-promoting agent such as phenols or ketones, e. g., linseed oil heated with CaO_2 and $ZnSO_4$ produces when the product is dissolved in turpentine a varnish too viscous for use, but a suitably reduced viscosity may be attained by adding about 4% of *o*-nitrophenol or benzoic acid to the modifying agents. The modified products may be used as softening agents with nitrocellulose. Products sol. in acetone or benzene are obtained by heating castor oil with Li_2CO_3 or Li sulfite to which *p*-cresol has been added. Cf. C. A. 24, 2910

Varnish. "CIRINE-WERKE" JOSEF LORFENZ & Co. Swiss 142,167, April 18, 1929. A weather- and acid-proof varnish is made by pulverizing a composition made from several varieties of wax, cold-emulsifying with a dry vegetable oil and mixing with a soln. of Co linoleate in dry vegetable oil and a wax solvent. Thus, bees-, carnauba-, mineral and shellac wax are ground up and emulsified with linseed oil. A paste of Co linoleate and linseed or wood oil and CaH_2 is then added.

Varnish bases. BAKELITE G. M. B. H. (Fritz Seebach, inventor) Ger. 517,445, Feb. 26, 1929. Phenol aldehyde condensation products capable of being hardened are combined with partly or completely oxidized drying oils. The materials may be combined in soln. and the solvent then evapd. The process may be effected in the presence of an O-carrier, and the oil may be oxidized *in situ*. Examples are given.

Lacquers. KAZUO TAKEMURA and KAZUO OTSU Ger. 517,349, Aug. 15, 1926. Nitrocellulose is dissolved in $AcOAm$, Et_2O and acetone, and the soln. is dild with $EtOH$ and Bu tartrate. $K_4Fe(CN)_6$ and powd. bamboo cane are then incorporated.

Oil lacs. I. G. FARBENFABRIK A.-G. (Karl Ott and Hanns Bernard, inventors), Ger. 514,914, Sept. 11, 1927. See Fr. 660,748 (C. A. 24, 252).

Non-inflammable nitrocellulose lacquer. MARKUS THAU, U. S. 1,793,726, Feb. 24, 1929. CCl_4 is added to a soln. of nitrocellulose in a solvent mixt. such as $BuOAc$ and turpentine.

Coating sheet materials such as wallboard with enamel. JOHN W. COWELL, U. S. 1,793,437, Feb. 17, 1929. Mech. features.

Coating metal foil. STANOLFABRIK BURGDOFF A. G. Brit. 338,340, Oct. 9, 1929. Sn or other foil is coated with a soln. of nitrocellulose in Et lactate and alc., an alc. shellac soln. and castor oil, with or without $AmOAc$.

Impregnating porous materials. ADOLF KANZOK, Swiss 141,109, Jan. 2, 1928. The materials are soaked in a liquid contg. a phenolic soln. of a substance which will not react with the phenol on forming synthetic resin, but which will promote the polymerization of the phenolic resin. Thus, wood is steeped in a mixt. of aldehyde- NH_2 , and cresol and heated to about 160° . The aldehyde and cresol form a synthetic resin which later polymerizes under the contact action of the NH_2 .

Films of cellulose esters or ethers, etc., suitable for tipping cigarets. CHEMISCHE FABRIK VON HEYDEN A.-G. Brit. 337,501, Feb. 15, 1929. Thin bands are formed comprising 2 continuous layers of cellulose esters or ethers such as cellulose acetate and ethyl or benzyl cellulose or the like, one layer contg. a pigment such as a metal powder while the other layer contains a material such as powd. magnesia, lithopone, powd. cellulose or cork which enhances its affinity for adhesives. Various details of manuf. are described.

Revivifying decolorizing clays containing rosin impurities. ROBERT C. PALMER (to Newport Co.) U. S. 1,794,537, March 3, 1929. Clay such as that which has been used for purifying rosin is treated with a mixt. of alc. and an org. solvent such as petroleum naphtha to dissolve the impurities in the clay, and the soln. thus formed is displaced from the clay with an org. alc. solvent such as a further quantity of petroleum naphtha.

Revivifying fuller's earth contaminated with rosin impurities. ROBERT C. PALMER, U. S. 1,794,539, March 3, 1929. Fuller's earth which has been used as a filter for purifying rosin is washed with a soln. of a rosin solvent such as petroleum naphtha and an org. revivifying liquid such as acetone having a high miscibility with the rosin solvent em-

proved and at least moderate miscibility with water at the temp. of the process (which may be about 65°). Mention is made of various other solvents which may be used.

Removing water from fuller's earth such as that used for purifying rosin. ROBERT C. PALMER and JOHN L. BURDA. U. S. 1,794,538, March 3. The material is treated with a substantially anhyd. liquid immiscible with water such as petroleum naphtha at a temp. (sensibly about 130-150°) above the normal initial b. p. of the liquid and under pressure, and the pressure is then reduced to vaporize part of the anhyd. liquid together with the water contained in the earth.

Casings of weighing apparatus formed of resinous material. C. F. M. VAN BERKEL. Brit. 338,424, Jan. 11, 1930. Casings of weighing app. are formed wholly or in part of molded natural or synthetic resin compn. which may comprise a filling of wood flour, asbestos or linen or other fabric. Various details of manuf. and construction are described.

Refining wood-resin. WILL SCHULTZ (to A. Schultze & Co.). U. S. 1,793,967, Feb. 24. Resin obtained from dead trees is distd. in a high vacuum (of 8 mm. Hg or less) in the absence of foreign gases.

Modifying the properties of resins. L. ATEX. Brit. 337,733, April 24, 1929. In a process such as is described in Brit. 321,724 (C. A. 24, 2904) for modifying the phys. properties of resins contg. high mol. org. acids, such as colophony or ester gum, the resins are heated under atm. or higher pressure with org. compds. of the O acids of S or their Cl derivs. such as aromatic sulfonic acids or chlorides. The material such as colophony or ester gum may, e. g., be heated at 280-300° in an open vessel with *p*-toluenesulfonyl chloride, 2,5-dichlorobenzene-sulfonic acid, *β*-naphthol-3,6,8-trisulfonic acid or sulfosalicylic acid. The products are either viscous oils or are brittle solids and may be used in varnishes. They may be subjected to a supplementary gas treatment or to irradiation. Brit. 337,734 relates to the modification of the phys. properties of fats, fatty oils or fatty acids by heating and dispersing in them modified isocolloids such as those prepd. by heating with amines (Brit. 337,732, following abstr.) or with compds. comprising an org. residue united to an inorg. acidic residue (such as benzenesulfonic acid) as described in Brit. 321,724 (C. A. 24, 2904) or with water insol. compds. as described in Brit. 321,725 (C. A. 24, 2904), or by treating an aq. soln. of the salts of amines such as benzidine or bisaniline with a soap soln. of high mol. org. acids. The products may be irradiated with ultra violet or x rays, treated with gases such as SO₂ or N₂, and may be vulcanized; they are suitable for varnish or rubber-substitute manuf.

Modifying isocolloids for varnish or rubber-substitute manufacture, etc. L. ATEX. Brit. 337,732, April 24, 1929. Materials such as l. seed or China wood oil or their fatty acids are heated to above 250° (or thickened oils or tung oil to above 200°) with 2-10% of a primary aromatic diamine or a monoamine of high mol. wt. or both, such as *m*-phenylenediamine, *p*-tolyl-nediamine, *β*-naphthylamine or *p,p'*-diamines such as *p,p'*-diaminodiphenylamine or *p,p'*-diaminodiphenylurea. Various other substances may be added and the products may be vulcanized with or without use of an accelerator or activator. Cf. C. A. 24, 2904.

Modifying natural and artificial resins. L. ATEX. Brit. 337,750, April 24, 1929. Various materials such as colophony or ester gum are heated under less than atm. pressure with metal free inorg. S compds. such as S chloride, chlorosulfonic acid, H₂SO₄ or sulfonyl or thionyl chloride (sensibly at a temp. of about 220-300°). The products thus formed are brown viscous oils or semi-solids suitable for use in varnishes and may be irradiated with ultra violet or x rays or mixed with ketones, phenols, etc. Cf. C. A. 24, 2904.

Synthetic resin from xylenols and furfural. EMIL E. NOVOTNY (to John S. Stokes). U. S. 1,793,715, Feb. 24. In making a potentially reactive resinous product, a mixt. of the ingredients is heated to above 100° and the water of soln. and of reaction is removed substantially as fast as it is vaporized while the reactive ingredients are returned to the zone of reaction, and a methylene-contg. hardening agent such as formaldehyde is combined with the reaction product. Cf. C. A. 24, 5173.

Synthetic resins. CARLETON ELLIS (to Ellis-Foster Co.). U. S. 1,793,310, Feb. 17. A chlorination product of toluene, chlorinated in the side chain and contg. benzotrichloride is caused to react on a phenol in the presence of a chloride of Fe or Al to produce a resinous product, and the latter is acted upon with about 10% of CH₂O. U. S. 1,793,311 relates to products similarly formed with the further addition of (CH₃)₂N₂, and U. S. 1,793,312 also relates to details of procedure of processes of the same general character.

Synthetic resins. BRITISH CELANESE, LTD. Brit. 338,002, Nov. 17, 1928. A toluenesulfonamide-aldehyde condensation product is further condensed with a toluene-

sulfonamide, and an addn. of 5-10% urea may be made to the initial materials, and the reaction may be effected in the presence of a catalyst of acid or alk. character or in the absence of a catalyst. The products are suitable for use in lacquers, plastic compns., etc., together with cellulose derivs., other synthetic resins, natural resins, solvents, plastifiers, modifiers, etc., various details and examples being given both of the prepn. and use of the products Cf *C A* 24, 2317.

Synthetic resins. BRITISH CELANESE, LTD Brit 238,024, Dec. 5, 1928. Resinous products prepd by condensing a toluenesulfonamide-aldehyde condensation product with a toluenesulfonamide, as described in Brit. 238,002 (preceding abstr.) are modified by further condensing with an aldehyde, *e g.*, a benzene-insol product formed as described in Brit 238,002 may be heated with formaldehyde soln under reflux at 120-150° to form a non resinous product which is converted into a benzene-sol. resin by heating at 130-150° in an open vessel

Synthetic resins. WILFRED S ROTHEN, STANLEY BLYTHEN and HENRY R. GILLSPIE Ger 517,251, Feb 16, 1926 Hydrophobe resins are prepd by warming urea or a urea deriv with an aldehyde, particularly CH_2O , for a short time, treating the primary condensation product so obtained with an org compd capable of condensing with the aldehyde in an acid medium, *e g.*, PhOH or thiourea, and then heating the mixt in the presence of free H ions until the resin ppts from the mixt. on cooling. The mixt. may be neutralized or rendered feebly alk before cooling Examples are given

Synthetic resins. I G FARBENFABR A-G Brit 238,109, May 9, 1929 Cyclic hydrocarbons contg at least one aromatic nucleus but having no olefinic side chains, such as benzene, naphthalene or tetrahydronaphthalene, or their alkylated, halogenated, nitrated, alkoxylated or carboxylic ester derivs., are condensed with diolefins such as 1,3-butadiene, isoprene or 2,3-dimethyl 1,3-butadiene, in the presence of anhyd. inorg halides which evolve heat on contact with water such as AlCl_3 , ZnCl_2 , FeCl_3 , B chloride or B fluoride, or of "ansolve" acids derived from such halides, such as complex compds of the halides of Zn, Ca or Mg with HOAc Products are obtained suitable for use in lacquers, etc., and several examples are given

Synthetic resins. I G FARBENFABR A-G (Friedrich Frick, inventor). Ger. 517,430, Dec. 29, 1928 Resins are prepd by condensing urea with Clf_2O in the presence of aryl ethers of glycols or polyglycols, or lower fatty acid esters of such ethers. About 0.5 mol of glycol deriv may be taken for each mol of urea. Examples are given Cf *C A* 25, 835

Synthetic resins. I G FARBENFABR A-G (Leo Rosenthal, inventor). Ger 517,477, May 1, 1928 Resins are prepd by treating crude solvent naphtha, alone or contg phenols or naphthols, with compds of BF_3 and fatty acids. The reaction may be effected by gradually adding about 3% of a BF_3 -fatty acid compd to crude solvent naphtha while stirring and preventing rise of temp above 70°. The products are colorless or nearly so, and are useful as varnish bases Examples are given

Apparatus for cutting thin sheets from a heated block of fully hardened artificial resin. HEALD A-G Brit 337,530, Sept. 13, 1929. Structural features.

Urea-formaldehyde resins. TOLEDO SCALE MFG Co Brit 337,357, April 27, 1929. Urea and Clf_2O are caused to react in aq soln to form an initial condensation product, the soln is concd., a condensation product of a "polybasic" acid and a polyhydric alc is then added and the soln is further concd. In the initial condensation, the pH may be adjusted to from 5 to 6 by adding triethanolamine, and after such condensation the pH may be increased to 7 or 8. Sufficient of the "polybasic acid"-polyhydric alc condensation product may be added to lower the pH to 3.5 to 5.0 and the mixt. may be dried by heating *in vacuo* and may be molded under pressure and cured by heating. Thiourea may be substituted for part of the urea Cf *C A* 24, 6041.

Linoleum or similar material. VICTOR SCHOLZ (to Atlas Agro Chemische Fabrik). U. S. 1,794,325, Feb 24 Oil such as linseed oil is preliminarily converted into a jelly by heating and blowing with air and is then subjected to a kneading process to give the product a fine crumbly, flaky, yellowish opaque character Cf *C A* 24, 2316

Decorating linoleum and similar materials. H. W. PRENTIS (to Armstrong Cork Co.). Brit. 338,205, Oct. 9, 1928 Various mech features of operation are described.

Forming color patterns on linoleum, felt, oil cloth, etc. J. C. MCCARTHY (to Armstrong Cork Co.). Brit. 337,924, Nov 21, 1928 Various details are given of procedure for successive color applications Cf *C A* 24, 2823

Floor covering. JULIAN T. BALDWIN (to Sandura Co.). U. S. 1,793,666, Feb. 24. A thin sheet material such as paper is placed on a supporting base such as a linoleum

or felt base with an intervening cementing material comprising a resin such as rosin, shellac or ester gum, a plasticizer for the resin such as diethyl phthalate and a "modifier" such as montan wax, rubber or oil and the product is provided with a decorative coat such as a pigmented resin compn and an overlying transparent wear resisting coating such as a shellac varnish contg a plasticizer. U S 1,793,667 relates to floor covering material prep'd by mixing body ingredients such as cork, wood flour, starch or rubber with a binder such as a resin and plasticizer and affixing the resulting mixt under pressure to a fibrous support such as textile fabric.

27-FATS, FATTY OILS, WAXES AND SOAPS

R. SCHERTEL

A new method for determining iron in fats and soaps. O. HAGEN. *Seifensieder-Ztg* 58, 110-2 (1931).—H detg Fe in mls by carefully ashing 10 g. and detg Fe in the HCl soln. of the ash colorimetrically with standard Fe soln. and KCNS. The following results are tabulated: colorless coconut oil 0.4-0.9, 1 A lard 0.6, 11 A lard 27.0, 1st pressing castor oil 0.9, tallow 4.5-36.0, coconut oil fatty acids 20.7 mg Fe per kg. of oil. Soaps are ashed in successive portions, a total of 5 g. is used and combustion is aided with a few crystals of NH_4NO_3 . The titration with KCNS gave for soaps a content of 2.3-25.2 mg Fe per kg. of soap. In a series of expts. for development of brown spots, the unchanged soaps contained 1.4-1.8 and the discolored soaps up to 9.0 mg Fe per kg. P. ESCHER.

Wool fat. P. MARTELL. *Pharm. Presse, Wiss.-Prakt. Hefi* 1931, 17-8.—A description of the secretion of wool fat and of its phys. and chem. properties in connection with the usual procedures followed in the production of the crude and refined brands. W. O. E.

Wool fat, its properties and composition. G. SALOMONE. *Boll. Icniera* 44, 199-201 (1930).—The properties and compn. of wool fat depend greatly on the way it is extd. The wax extd. with petr. ether has a thicker consistency and is richer in hydroxy acids than that extd. with trichloroethylene. It is brown or brown yellow, has no disagreeable odor, is little sol. in 95% alc., much more in ether, in benzene, in chloroform, trichloroethylene, CS_2 , and is dextrorotatory. It has $d_{20} 0.941-0.945$, $n_D^{20} 1.38-1.42$, sapon. no. 98-115, i. no. 17-29, unsaponifiable 30-52%, fatty acids in 42%, mol. wt. of acids 327.3. The fat extd. with petr. ether or acetone has little acidity, while that obtained from scouring baths has 20-25% of free fatty acids calcd. as oleic acid. Wool fat is difficultly saponif., it contains lactones that are transformed gradually by the alkali into acids and their salts. There is a little myristic acid and a large amt. of palmitic, stearic, carnaubic and cerotic acids. Small quantities of cetyl, carnaubyl and ceryl alcs. are present with cholesterol, ischolesterol and oxycholesterol, in a free state in limited amount, and for the greater part esterified with the fatty acids. Such alcs. represent the unsaponifiable portion that can reach 50% of wool fat. The esters predominating in wool fat are principally palmitate, stearate and cerotate of cholesterol and ischolesterol. The identification of wool fat is based on the presence of cholesterol. Impure wool fat, if exposed for some time to air, becomes brown, and acquires a less agreeable odor, a greater consistency, a higher refractive index, a lower i. no., a less complete soly. in petr. ether and a higher mol. wt. by the decompn. of the cholesterol esters that are first hydrolyzed and then oxidized. R. SANSONE.

Freezing-point method for the examination of cacao butter. A. G. AVENT. *Analyst* 56, 180 (1931).—In the paper published on the classification of chocolate fats by the f. p. method (C A 24, 4946) mention should have been made of the assistance obtained from H. R. Jensen in devising the standard solidification test. W. T. H.

Extracted cacao butter—imitation or adulterated. W. NORMANN. *Allgem. Öl-Fett-Ztg* 27, 259-61 (1930).—The inconsistency of German rulings on the subject of extd. edible fats is discussed. W. F. BOLLENS.

Measurement of rate of formation of oxidative decomposition products in fats and oils. D. P. GRETTIE AND R. C. NEWTON. *Ind. Eng. Chem., Anal. Ed.* 3, 171-3 (1931).—The method used is a modification of the methods of Bailey and Issoglio. The measurement of the volatile products is taken as a measurement of the degree of spoilage of the fat, and is performed by titrating against acid KMnO_4 ; the products carried away in a given time by a const. stream of air pass over the fat which is dispersed on filter paper and maintained at a specified elevated temp. The stream of air is drawn over the fat and then through a standard soln. of acid KMnO_4 maintained at 25°. This

soln. is then titrated against oxalic acid to det. how much was used by the decompn. products condensed in it. E SCHERUBEL

Causes of deterioration of fats and oils. W. L. DAVIES *Oil and Fat Ind.* 7, 427-31, 453-7(1930).—The 2 main causes of deterioration are oxidation and the agency of microorganisms. O absorption is not rapid until after the induction period has been overcome. This means the time required by the system to undergo activation. The conditions likely to catalyze the subsequent oxidation of fats have also the effect of shortening the induction period, e. g., heat, light, acidity, other impurities and metallic catalysts. Lipoids in fats are oxidized first, thus in butter the fishy flavor is due to oxidation of lecithin forming trimethylamine. Fats which contain lipase and protein decompn. products also shorten the induction period. The tests used for following the course of oxidation such as the Kreis test and the extn. of the oxidizability values of the steam volatile or water sol. products of oxidation have been found valueless in detecting incipient deterioration. When bacteria are present in fat no autocatalytic oxidation occurs as the organisms use all available O in the fat. Molds cause more deterioration in fats than bacteria. They are of 2 classes: those which preserve the solid consistency of fat and those which have strong lipolytic properties and liberate fatty acid. Very little is known of the effect of yeasts on fat. E SCHERUBEL

Determination of rancidity in oils and fats. ALAN TAFTEL AND CECIL REVIS. *J. Soc. Chem. Ind.* 50, 57-91T(1931).—On account of emulsification caused by ether in the Kreis reagent and the presence of peroxides which tend to oxidize phloroglucinol, the Kreis test was modified by substituting alc. for ether and using 10 drops of a 5% soln. of phloroglucinol in alc. Oils which have been air blown or have become rancid at moderate temp. contain peroxides which are reduced by HI and oils becoming rancid around 120° contain more peroxides together with other forms of combined O which offer more resistance to reduction by HI. Oils air blown at 170° contain only difficultly reducible peroxides. The amts. of each type of oxide present in an oil can be detd. For oils becoming rancid at moderate temp. add 10 g. of oil to 40 cc. glacial AcOH and then 2 cc. of 50% KI soln. Shake the mixt. for 2 min., hold for 2 min. and again shake for 2 min. and then pour into 100 cc. H₂O, rinse with 20 cc. and titrate with 0.1 N Na S₂O₄ and starch. Two g. of BaI₂ may be used instead of the KI soln. For oils which have become rancid by exposure to air at high temp. the proportions of the reagents must be found by expts. Ten to 20 g. of BaI₂ and 100 cc. of AcOH are suitable for 1.25 to 2.5 g. of oil. For the samples examd. the Kreis reaction and the I or peroxide figure showed a close correlation, a strong reaction of the former accompanying a high peroxide figure. While the peroxide methods have not stood the test of long experience they are of apparently general application to all types of rancidity. Definite and reproducible figures are obtained by the methods which distinguish between oxidation in ordinary rancid oils and that which occurs in oils blown at high temps. E S

Three new oils. ERICH STROCK. *Farben-Ztg.* 36, 830(1931).—The analytical consts. of 3 new oils were

	<i>Morqualis tomentosa</i> Beeth.	<i>Oenocarpus</i>	<i>Udio</i>
sp. gr.	0.8864 (at 16°)	0.9244 (at 15°)	0.9241 (at 15°)
solidification pt.	-14.5°	7.5°	7.9°
n _D	1.4919 (at 30°)	1.4700 (at 15°)	1.4699 (at 15°)
sapon no.	196	190	198.9
acid no.	18.8	1.33	66.9
I no. (Wijs)	81.5	80.6	98.8
unsaponifiable %	7.5	0.84	
resins			19.4%

G. G. SWARD

Some titer points of mixed fatty acids. I. Mixtures of commercial oils, fats and fatty acids. GEORGE W. JENNINGS. *Ind. Eng. Chem.* 23, 413-5(1931).—Mixed fatty acids of tallow, corn, olive, coconut and cottonseed oils, garbage, grease and red oil were prepd. and mixed with each other in the desired proportion just before detn. of the titer. The results are reported in the form of curves, plotted with the titer points as ordinates and the % compn. of the mixed fatty acids as abscissas. The work is to continue in order to explain the cause of the apparently erratic results obtained in some cases. E SCHERUBEL

Intramolecular rearrangement of the ester of simple unsaturated fatty acids during hydrogenation. A. STEGER AND H. W. SCHEFFERS. *Chem. Umschau Fette, Oele, Wachse Harze* 38, 45-53(1931).—S and S used specially prepd. pure ethyl, glycol and glycerol esters, etc., for a partial and complete hydrogenation with Ni on kieselguhr—

reduced at fixed temps. and with Pd on C according to Mannich and Thiele (C. A. 10, 218). They exam'd the influence of Ni reduction temps., of the hardening temps., and of the amt. of catalyst upon the quality of the product. The results of the detns. of the I no. of the liquid and solid acids, % solid acids and % linoleic acid are tabulated and plotted in graphs. Results. Solid isoleic acid is formed soon after hydrogenation. Higher temp. increases it and a max. is reached at 50-60° I no., linoleic acid disappears rapidly and becomes oleic acid especially at higher temps.; no stearic acid is formed until only a few percents of linoleic acid remain. Oleic esters have their oleic acid in part converted into isoleic acid which reaches a max. and then diminishes, disappearing entirely around I no. 20. Ft esters form more iso-acid than glycerol esters. Solid elaidic ester forms liquid oleic acid, more of it with Ft ester than with glycerides. Changes in the reduction temp. of Ni are not reflected in the phys. properties of the product. An increase in the amt. of catalyst increases the isoleic acid but slightly. A Pd catalyst forms more iso-acid than Ni, but the type of carrier (Lieselguhr and C) does not influence the results. P. FISCHER

Detection of sesame oil in solid mixtures of fats. R. LUCCENTINI *Ind. chim. minerali e grassi* 10, 146-7 (1930).—If less than 0.5% of sesame oil is present ext. 50 g. of the melted fat for 2 hrs. with 100 cc. 95% alc., using a Marino extractor cool to rpt. the excess of dissolved fat, filter and evaporate the alc. on the water bath. The residue contains, with a little fat, the greater part of the phenolic substance that produces the color reaction for sesame oil. After cooling, dissolve the residue in 5 cc. of petr. ether, bring it into a test tube contg. 2 drops of a furfural soln. in alc., add 5 cc. of concd. HCl and agitate several times softly. In the presence, even of very small quantities of sesame oil, a pink red color is produced. R. SANONE

Vitamins and palm oil in margarine. ALBERT K. ERSTEIN, *Oil and Fat Ind.* 8, 107, 109 (1931).—Tests made on samples of palm oil made from selected raw material and which were refined and deodorized by careful methods were found to respond to the Drummond color test for vitamin A equiv. to about 60 units per g. This is a higher test than that given by the av. butter. E. SCHWARTZ

Present state of the olive-oil industry in Italy. HUMBERTO MORINI *Chim. e ind.* 7, 232-5 (1930). S. L. B. LITERTON

Analysis of olive oil in Wood's light. G. LOEW, *Ind. chim. minerali e grassi* 10, 140-2 (1930), cf. C. A. 24, 1532.—Tests showed that the olive fluorescence of olive oil is due principally to the presence of vitamin C or to a substance that accompanies vitamin C. R. SANONE

Utilization of the aqueous extracts from olives. RICCARDO CITSA, ANGELO MANGONI and DOMENICO MASSIMO *Industria Chimica* 5, 1451-4 (1930).—The aq. ext. obtained during the manuf. of olive oil, previously rejected, has been found to undergo spontaneous fermentation, a max. of 3% EtOH being obtainable. In addn. 1/10% of oil can be recovered by centrifuging, as well as 2% of KCl. A. W. CONTI

The processing of the fish oils. G. HINARD *Boll. ufficiale staz. sper. ind. pell. mat. concianti* 8, 161-7 (1930).—The properties and uses of fish oils and of products obtained from fish oils are discussed and the methods involved in the prepn. of the products are described. G. SCHWACH

The behavior of bleaching powders in acid oils and attempts at removing free acid by distillation. LEV PICK, *Allgem. Öl-Fettig* 27, 291-4 (1930).—Bleaching earths and carbons, when mixed at 103° with peanut oil or coconut oil, contg. free fatty acids were found partially to remove the free acidity. This was found to be true for acid earths as well as for neutral earths, when amts. of 10-50% were used. This removal of free acid is thought to be due to the adsorption of the free carboxyl groups to the surfaces of the bleaching powder. This interferes with the adsorption of pigments, and thus accounts for the lower bleaching power of earths and C on acid oils. Attempts to remove the free acidity from peanut oil by steam distn. at atm. pressure were only partially successful. The best results were obtained at 287°, the free acidity being reduced from 3.5 to 2.0%. During the distn. 2.7% of fatty acids was distd. off, showing that hydrolysis had taken place. When this procedure was used on coconut oil, the free acidity was reduced from 2.2 to 0.2%. The distn. temp. was 305°. When 3% of coconut oil was added to peanut oil (free acid 4.5%) and subjected to the above process, at 300°, the free acid was reduced to 0.4%. This difference in behavior of the 2 oils is said to be that, at temps. of 280° to 300°, the higher free acids replace the lower acids of coconut oil, which at this temp. are readily distd. leaving a practically neutral oil. Attempts were also made to remove the free acid by distn. with a volatile solvent. An app. is described in which the oil dissolved in benzene (2:1) is subjected to a continuous distn. under a reduced pressure of about 100 mm. Hg. The benzene soln. is

drawn through a long Al tube submerged in an oil bath at 300°, and thence into a flask submerged in an oil bath at 250°. The neutral oil collects in this flask, while the fatty acid and benzene vapors are carried through a condenser into a receiver. This procedure reduced the free acid of peanut oil from 4.0 to 0.1%, and that of coconut oil from 1.1 to 0.1%. In another sample of coconut oil the free acid was reduced from 13.6 to 0.7%. W. F. BOLLINS

Substitution of steam for carbon dioxide in the hydrogenation of oils. V. YASUCHIKO. *Maslobonna Zhirorv Delo* 1929, No. 5, 21-4, *Chimie & Industrie* 23, 427-8. —Before removal of the catalyst after hydrogenation CO_2 is passed through the hydrogenation vessel. Passing steam at about 120-140° (to prevent condensation) was found equally satisfactory. A. PAPINIAU-COUTURE

Preparation of nickel catalyst for the hydrogenation of oils. A. HAO. *Maslobonna Zhirorv Delo* 1929, No. 5, 32-4, *Chimie & Industrie* 23, 428. —In the regeneration of Ni catalyst the Fe^{++} being present in the Fe^{++} state ppt. as FeCO_3 , along with the NiCO_3 instead of with the Al and Zn. This is overcome by aerating the soln before pptn. to convert Fe^{++} into Fe^{+++} . A more active catalyst is obtained by pptg. the Ni by adding the Na_2SO_4 soln. to the mixt. of Kieselguhr and Na_2CO_3 soln., instead of vice versa. A. PAPINIAU-COUTURE

The effect of supports on the catalytic activity of nickel. CHARLES R. GLASS AND LOUIS KAULNINKO. *Trans. Elektrochem. Soc.* 59, preprint (1931). —Those support materials which inhibit the catalytic activity of Pt and Pd have been found to effect Ni similarly. The behavior of the supports with relation to the hydrogenation of cottonseed oil by Ni can be classified as follows: Reducible heavy metal oxides completely inhibit the activity of Ni. The salts of those metals that decompose on heating in a stream of H_2 completely inhibit activity. Substances with corrosive properties retard but do not completely inhibit activity. The salts in every case have been found to yield more active supports than the corresponding oxides. The more inert the support becomes toward Ni the greater is its tendency to enhance activity. The less inert the support becomes toward Ni the greater is its tendency to retard activity. The oxides of Sb, Al, Cr, Mn, Ti, Ce, U and Mg and the phosphates and borates of Al, Cr, Mn, Ce, U, Mg, Ca, Sr, Ba and Li proved very beneficial as supports for Ni in hydrogenation. C. G. F.

The bleaching of waxes. J. DAVISON. *Allgem. Öl. Fettz.* 27, 261-2, 270-7, 312-3 (1930). —A review. W. F. BOLLINS

Sodium chloride determination in soaps. STÄNCHUN. *Seifensieder-Ztg.* 58, 112-3 (1931). —S combines the NaCl detn. with that of the total alkali and the fatty acids, using Volhard's KCNS method for the NaCl detn. Decompose 5 g. with known HNO_3 , shake out with ether, wash the ether soln. for fatty acid detn. To the HNO_3 soln. add 5 cc. ether per 100 cc. soln., shake out, add 5-10 cc. 0.1 N AgNO_3 , shake well and set aside for 5 min. In the dark; add Fe indicator in HNO_3 soln. and titrate back with KCNS to a light rose tint. P. ECHER

Properties of soap solutions. I. Estimation of detergent capacity. B. TRYUNNIKOV. *Allgem. Öl. Fettz.* 27, 211-3 (1930). —The ideal way of expressing detergent capacity would be to express it in terms of the energy, which a known soln. of known concn. and temp. exercises toward the overcoming of the attraction between the substrate and soil, and toward the suspension of the soil in the soln. This energy would be measured by phys. and colloidal chem. methods, rather than by the highly standardized exptl. washing tests. II. The influence of saponin on soap solutions. B. TRYUNNIKOV, N. KASSYANOVA AND R. GYREMAN. *Ibid.* 273-6, 291-8 (1930). —The influence of the addition of successive portions of saponin soln. to soap solns. is divided into 3 stages. In the 1st stage the lathering power is reduced until it is 0. During the 2nd stage the lathering power remains at 0 while during the 3rd stage lather is again produced. Curves showing the influence of addition of saponin on the lathering properties of Na soaps of various fatty acids and of rosin are given. Tables are also given showing the influence of addition of saponin to these soaps on the lather vol. (cf. C. A. 24, 3668), the property of carrying pptd. MnO_2 through a filter, and the surface tension. It is supposed that the saponin is adsorbed by the soap, and that the adsorbing power of the soap is not completely satisfied until the beginning of the 3rd stage. This may be numerically expressed by the amt. of saponin required just to bring forth a lather after having destroyed it, expressed in multiples of the wt. of soap present in the soap soln. The no. is called the saponin no. Saponins of different origins were found to give somewhat different results, and as yet no method for the evaluation of the various saponins has been developed. W. F. BOLLINS

Evaluation of detergency. JAMES G. VAIL. *Soap* 7, No. 3, 20-30 (1931). —The

trend in soiling tests has been in favor of simple pigments, carbon black and umber being preferred. Two methods of applying pigments to water give fairly even colors. In one 5 g. of umber in 100 ml. of water is used with 5 x 12 in. of cloth which is passed through a clothes wringer and the treatment repeated until the cloth color is constant. The second method involves the use of the Launderometer, a machine with a round bottom metal tank supported in an open frame and having a shaft bearing 16 threaded posts. By means of clips and wing nuts 20 jars may be attached. A motor drives the shaft at 40 r. p. m. Heat is applied by a gas burner underneath. A cloth bag containing 50 No. 1 metal balls is rotated for 1 hr. at 60° in a suspension of 1 g. of umber in 100 ml. of water. This machine is also used for the washing tests. For judging results a color chart was devised but results did not agree and another plan was to print the color in stripes across the cloth. Washing was to be continued until the marks were invisible but it was difficult to judge when they disappeared. At present no procedure has been adopted. F. SCHERBEL.

Fatty acids of the resins of jack pine (PHILLIPS) 23. The oxidation of the fat used in oiled wood. SALOMONE 25. Floating factories [for whale oil] (HAUPTVOGEL) 15. Calculating viscosities and flash point in compounded oils (KADNER) 2. Report of the oil and fat commission (AUESSACH) 29. The higher alcohols of the paraffin series [as waxes] (SCHRAUTH) 10. Wax emulsions [for soaps] (Brit. pat. 338,176) 18. Solvents for fats and waxes (Ger. pat. 517,997) 13. Thickening oils (Brit. pat. 337,639) 22. Method and plant for separating waxy and fatty substances from oils (Ger. pat. 517,198) 22. Rotary impeller agitator device and filter for treating oils (U. S. pat. 1,794,916) 1. Apparatus for extracting oils with liquid solvents (U. S. pat. 1,794,874) 1. Unactivated 'prefiltering medium' for treating oils (U. S. pat. 1,794,862) 1. Distillation of volatile substances [in refining oils and fats] (Ger. pat. 514,749) 13. Apparatus and procedure for spraying desiccation of soap (U. S. pat. 1,794,978) 1.

Fat separators. WILHELM STEINMANN. Swiss 142,451, Nov. 2, 1929. A fat separator for a plant for working up animal waste, carcasses, etc., is described.

Extracting oils and grease with solvents. IMPERIAL CHEMICAL INDUSTRIES, LTD., and F. F. ROGERS. Brit. 337,770, Aug. 1, 1929. Various details of app. are described for use in a process, such as described in Brit. 278,691 (C. A. 22, 2431). A condenser and the means for heating the app. are interconnected so that the cooling of the condenser ceases when the heat is shut off.

Organic peroxide for oxidizing or decoloring. HENRY B. FASER (to S. G. Turnbull). U. S. 1,793,917, Feb. 24. Substitution of O for Cl in material comprising an aliphatic dicarboxylic acid chloride and an aromatic monocarboxylic acid chloride (such as the reaction products from heating fumaric acid with benzotrichloride at 140-170°) is effected by treatment with H₂O₂ in the presence of an alk. substance such as Na₂CO₃. The product may be used for oxidizing or decolorizing oils, cereals, cereal products, etc.

Leaching materials such as in extracting oil from seeds. JESSE M. COCHRAN. U. S. 1,793,465, Feb. 24. Various details of app. and procedure for countercurrent extr. are described.

Obtaining oil, protein material, starch and cellulose from seeds. LOUIS F. DAVID and GEORGES FELIZAT. U. S. 1,794,195, Feb. 24. Seeds are crushed and the crushed mass is mixed with a slightly alk. aq. soln. such as Na₂CO₃ to dissolve protein material, swell the cellulose and suspend the starch and oils (the latter being in emulsion), the cellulose is sep'd. by screening, the screened liquid is centrifuged at slow speed to suspend starch, and the liquid is then centrifuged at higher speed to break the emulsion and sep. the oil. The liquid from the second centrifuging is treated with acid to ppt. protein material.

Apparatus for separating solvent from extracted seeds. KARL LÜFFL. Ger. 514,915, Feb. 6, 1929. The greater part of the solvent is recovered by blowing non-condensable gas through the seeds, and the remainder by squirting the seed through a nozzle into a chamber.

Scouring composition. W. J. A. HUYZER. Brit. 337,355, July 26, 1929. Soap powder is mixed with 1-3 times its quantity of natural magnesite ground to approx. 50 mesh per linear in.

Dry soap. ALBERT RUF. Swiss 142,452, May 13, 1929. A good ordinary soap is dissolved in boiling water. Calcined soda and a water fixing stabilizing agent are added. The whole is reboiled and spirit added during violent agitation. On cooling, the whole is worked up to a creamy mass. In the example palm-olive soap forms the starting material.

Apparatus for molding soap. L. H. NIKINA Brit 377,430, May 2, 1929 Structural features

Cold press for soap, with heat-absorbing plates. AUGUST JACONI A-G Ger. 614,626, Dec 20, 1929

28 SUGAR, STARCH AND GUMS

J. K. HARR

A hundred years ago. X. R. O. V. LIPMANN Deut Zuckerind 56, 18 9(1931); cf C A 24, 1530 —A history of best sugar J. P. LINTH

Cuba's five-year plan. EARL L. SYMON Intern Sugar J 32, 612-4(1930) —A description of the operation of the governmental control method for sugar production in Cuba which became effective Nov 15, 1930 W. L. OWEN

Technical notes from Usina Cuscu Pernambuco. N. E. LAMONT Intern Sugar J. 32, 629-31(1930) —The use of "Sannaphos" in clarification applied at the rate of 200 g per ton of cane at the filtering and sulfuring tanks results in an improved decantation and an increased yield of sugar, while the massecuites are free working and the washed sugars are much more easily filtered Expts made on the purging of second massecuites with warm oil exhausted molasses showed that the purity of the resultant molasses was on an av. 1% lower than without this aid and the actual time gained in curing each charge was reduced 20% Expts on the influence of pH on alc recovery showed that the closer the initial pH to that of the final fermented wash, the higher is the efficiency of yield An increase of 18% in yield was obtained by introducing the fresh wort on a looking of sour wash W. L. OWEN

Factory results in Java during the last seven years. R. J. PRINSEN GERRITSON Intern Sugar J 32, 631(1930) W. L. OWEN

Dilution in relation to comparative purities. NORM DEPPER Intern Sugar J. 32, 610-7(1930) —D discusses LINTH's article (cf C A 24, 6023) and replies to the exception taken by the latter to the statement that "to be strictly comparable, gravity solids deins should be made in solns of the same nonsugar concn" Taking the formula sugar recovered $S(S-M)/J(S-M)$, where S , J and M are the purities of the sugar recovered, the original impure material, and the residue or molasses, resp, D demonstrates that an accurate Brx or gravity solids balance can also be made where nonsugars are deid in equal concn The identity of the values obtained from the use of alcoholide and gravity purities shows what is meant by the word "comparable" in the official Direction of the Hawaiian Sugar Technologists W. L. OWEN

Cylindrical or centrifugal pumps (in the sugar industry)? V. SZAVASKY, Z. Zuckerind Technol Rep. 55, 127 8(1930) —A discussion S states that centrifugal pumps are suitable even for mud flows J. P. LINTH

An improvement in the testing of press mud for sugar. P. BRUCKNER AND TH. BERTHAUPT Ceur Zuckerind 38, 13 9(1930) The authors find that foaming and dark colored filtrates are avoided by substituting a soln contg 8% NH_4NO_3 and 2% $Al(NO_3)_3$ for the 10% NH_4NO_3 soln in the method of the Institut für Zucker-Industrie. Numerous expts indicate that no errors are introduced by the change J. P. L.

The calculation of the amount of massecuite in a white sugar factory. B. DIMIN. Deut Zuckerind 55, 1031-2, 1080-1(1930) A math discussion J. P. LINTH

Charging of sugar in centrifuging. W. C. HERRING Arch Zuckerind, 38, 250 (1930). —Sugar which was gray was found under the microscope to be permeated with particles of C, because of the fact that it had been washed with superheated steam which had been produced from boiler water contg sugar R. C. A.

Precipitation of lime by sulfuric acid in sugar solutions. EMILIE SAHLARD Compt rend. 102, 178 80(1931). —This study was undertaken to decide whether SO_2 added to limed juices combines first with the lime, with the alkalis, or with both at the same time. Solns, contg 10 and 20%, resp, of sugar were prepd. One g CaO and increasing quantities of $NaOH$ were added to each. Then SO_2 , amounting to less than its CaO equiv., was passed into the solns The alk., and the % SO_2 in the filtrates were deid. It was found that the SO_2 combines with both the CaO and the $NaOH$ As the quantity of $NaOH$ increases, more SO_2 is taken up by it At equal alk., more SO_2 combines with the CaO if the soln contains a large % of sucrose. I. W. ZERBAN

The use of the refractometer in cane-seedling selection work. N. CRAIG, Intern. Sugar J. 33, 14-18(1931) —The refractometer can be used to advantage in cane seedling work, the results obtained being very reliable. The method of sampling is rapid, re-

liable and easy of application. Whenever possible it is advantageous to take 6 borings from each stool, although the analysis of 3 borings is sufficient to give a fairly good indication of the sugar producing capacity of a cane seedling. The final choice of a seedling should be decided by the weights of the cane in the stool, and the richness in sucrose, and of these two the latter, which shows no correlation with the wt., is the more difficult to det.

W. L. OWEN

Recent investigations into sugar-beet problems. 1. Storage of sugar beet. G. R. CLARK, L. F. NEWMAN and A. W. LING. *J. Ministry Agr.* 36, 870-80(1930); *J. C. A.* 24, 2629.—Beets were stored in various types of clamps during one season, in the next season only the mangold type of clamp was used, with free circulation of air. Samples were removed at stated intervals, and the following detns made: loss in wt. dry substance in pulp, and % sucrose by Sachs-Le Docte method, the latter was checked by other methods. The results are shown in 8 tables. Conclusions: Under adverse climatic conditions beets lose both moisture and dry substance. Beets crowned before storage may seriously deteriorate. Secondary growth may start in imperfectly cut or in uncrowned beets. The net result is a decrease in the com. sugar yield. With normal weather conditions the loss is very small, and in favorable seasons beets can be stored without serious damage.

J. W. ZIEBMAN

Increasing the daily production or lengthening the (beet sugar) campaign. P. HIRSCHFELDER. *Deut. Zuckerind.* 56, 41-2(1931).—A discussion. If, points out 3 reasons in favor of a longer campaign; the low price of sugar, high interest rates and good storage facilities for beets. Reply. E. TROJE. *Ibid.* 95-6. J. P. L.

Comparative determination of the marc content of fresh and stored (sugar) beets by the methods of Claassen and of Thielepape and Meier. MATSCHKE. *Centr. Zuckerind.* 38, 1112(1930).—M describes and discusses his methods. The exptl. data show agreement within 0.10% by the 2 methods in most cases. The greatest discrepancy found was Claassen 5.31%, T and M 4.76%.

J. P. LEITE

The determination of the natural alkalinity (in sugar-beet juice): the method of the Institut für Zucker-Industrie or that of Düwell and Solon. K. DOWNING. *Deut. Zuckerind.* 55, 1051-82(1930).—D gives comparative results for the 2 methods and concludes that the D and S method is sufficiently accurate for factory control when only natural residual alk., and residual lime are to be detd., provided NH_3 is boiled out before carbonation, but recommends an occasional check by the method of the Institut.

J. F. LEITE

The sugar losses of beet-sugar manufacture. II CLAASSEN. *Intern. Sugar J.* 32, 560-1, 620-9(1930), *J. C. A.* 24, 2312.—The total losses in the usual method of working in beet-sugar manuf. amounts to 1.2-1.5% of the roots when all calens are carried out by correct methods. The correct estd. of the wt. of roots entering the factory is of prime importance. Through faulty detns this error may have a value of from -0.05 to +0.10. When returning the best wastes (tails) into the process, the quantity of sugar so introduced should be calcd. Because of inadequate sampling too low figures are usually found for the detd. loss, while with proper sampling they amount to from 0.4 to 0.6% of the roots. Undetd. losses, or losses undeterminable according to the usual methods, reach up to 0.75-0.90% of the roots, and are composed of apparent losses due to incorrect tares, increase of polarization of the beet slices through drying, polarization increases from the action of lime, too low results from the sugar content of the scums, sugar losses from decompn., and mech. sugar losses. Hence of the 0.75-0.90% undeterminable losses from 0.3 to 0.4% can be explained, while the remainder cannot be accounted for.

W. L. OWEN

Action of different decolorizing products in beet-sugar refining. C. GROSS. *Ind. Saccarif. Ital.* 22, 489-94(1929).—When decolorizing beet sirups with animal charcoal an excess of alkali is not beneficial, but increases the ash of the sugar. As the absorption of good charcoals is only 0.002% alk. between 0.01 and 0.015% is a sufficient protection against inversion during a first and second refining. 0.3% CaO does not cause sirups contg. a min. of invert sugar to become strongly colored, even when these are kept for more than 1 hr. at 90°. It often causes a slight ppt. Turbidity takes away a part or all the decolorizing power of the animal C. Only a limited portion of the color is removed by the animal C, an excess causing no further color removal. 0.035% gives good results. There was no improvement when the amt. was doubled. The addn. of hyposulfite always gave satisfaction. In reasonable amt., no inversion took place in slightly alk. massecuites. Fifty g. of hyposulfite is effective with 14 tons of first refining massecuites, while 14 tons of 2nd refining massecuites require 100 g. Most of the hyposulfite decoloration resists subsequent oxidation. R. S.

The question of the applicability of the conductimetric method for the official

determination of ash in raw (beet) sugars. J. PUCHERNA. *Z. Zuckerind. czechoslov. Rep.* 55, 205-6(1931). cf. *C. A.* 25, 1400—A discussion with statistical data. J. F. LEET

The conductometric ash apparatus (for beet sugar) in the practical field. Q. J. LIČKA. *Z. Zuckerind. czechoslov. Rep.* 55, 227-8(1931)—A discussion with status from 3 campaigns. J. prefers this method. J. F. LEET

Measurement of turbidity with a spectrophotometer, with especial reference sugarhouse products. R. T. BALCH. *Ind. Eng. Chem., Anal. Ed.* 3, 124-6(1931). Previous methods, none of which has come into extended use, are reviewed. Present method is based on a direct spectrophotometric comparison: the transmittance is determined with the turbid solution in one cell, and a filtered portion of the same solution in another. Filtration through paper with the aid of a slow filtering type of kieselguhr satisfactory. Kieselguhr does not cause true adsorption. Solids and semisolid products are diluted to approx. 50 Brix; any similar products to be compared should be diluted always to the same concentration. The transmittance measurements may be made at any convenient wave length, such as 560 mμ, but it was observed that the transmittance at the red end of the spectrum are higher than those at the blue end. Beer's law does not hold exactly, but the approximation is close enough to permit expression of turbidity in terms of $-\log I$. Expts. with bentonite suspensions showed that equivalent turbidity concentrates in a water-white medium and in one colored strongly with carmalum give the same transmittance readings. The results of turbidity measurements on a number of granulated and raw sugars are given in a table, and it is also shown that the turbidities found are roughly proportional to the % of suspended matter obtained by weighing. F. W. ZERBA

Working of molasses in the experimental molasses station at the sugar plant Galesno. J. ELIKS POLAK. *Przemysł Chem.* 15, 37-43(1931)—This work is basically on the sulfite fermentation of molasses. The apparatus and processes are discussed. The experimentation has reached a semi-plant stage in its development and includes fermentation and distillation of alcohol and acid oxidation of the aldehyde to HIOAc , oxidation of the sulfite to sulfate and distillation and purification of glycerol. A. C. ZACHIL

Acetone compounds of the sugars and their derivatives (OHLE, LICHTENSTEIN). Sugar-beet root rot control (YOUNG) 15. The use of excess molasses (CLAASSEN). The use of sugar for motor fuel (SANOERA, ZEMAN) 21. Report of the most important work in the field of pure sugar chemistry published during the first half of the year 1930 (v. LIPPMANN) 10. Unactivated "pre-filtering medium" for treating sugar solutions (U. S. pat. 1,794,862) 1. Rotary impeller agitating device and filter for treating sugar solutions (U. S. pat. 1,794,916) 1. Obtaining starch from seeds (U. S. pat. 1,794,105)

Sugar. RAFFINERIE TIRLEMONTAISE SOC. ANON. Ger. 514,713, Mar. 6, 1931. Air currents are prevented from circulating through the sugar crystals separated in a centrifuge, to prevent air drying of mother liquor adhering to the crystals before complete separation.

Sugars from cellulosic material. ROBERT GRIESBACH and ERNST KOCH (to I. Farbenind. A.-G.). U. S. 1,793,084, Feb. 17. Various details are given of a process involving treatment of cellulosic material such as pine sawdust with a mixture of H_2PO_4 and concentrated HCl .

Filter suitable for sugar solutions, etc. JOHN J. ARMSTRONG. U. S. 1,792,717, Feb. 17. Structural features.

Rotary filtering apparatus for filtering sugar solutions, etc. JOHN J. NAUG. U. S. 1,793,289, Feb. 17. Structural features.

Crystallization apparatus for sugar, etc. FERNAND LAFBUILLE. Ger. 517,117, May 3, 1928. See Brit. 301,453 (*C. A.* 23, 4095)

29—LEATHER AND GLUE

ALLEN ROGERS

Bibliography of papers on the physical properties of leather. III. C. H. SMITH. *J. Intern. Soc. Leather Trades Chem.* 15, 61-7(1931); cf. *C. A.* 24, 4951, 5178.—References are classified under: gaseous permeability (13 titles), behavior to bending forces (30); cracking of grain (46); pitting, break (14), thinness (5); resistance

leat (17) behavior on compression (6), hardness (3), resilience (3), other phys. properties (4 classes—6 titles) H B MERRILL

Report of the International Commission on Leather Analysis. R FARADAY LYNN. *J Intern Soc Leather Trades Chem* 15, 6-12(1931)—Current problems and recent advances are discussed. Topics include the detn of H_2O , ash, fat and acidity in vegetable tanned, and of basicity of Cr salt, fat, H_2O -sol matter and N in Cr tanned leathers. The theoretical basis of the Atkin-Thompson method (C. A. 23, 4813) for acidity of vegetable tanned leather—namely that the leather contains free amino groups that enable it to function like a gel with respect to acid—is questioned. Infusions of leather in H_2O and in 0.1 N KCl do not differ appreciably in pH value. Failure of petr ether completely to ext sulfonated oils and Al soaps is pointed out. H B M.

[Methods employed in] Chinese leather centers. H A. BRAUNSTEIN. *Hide and Leather* 81, No. 4, 30 (1931). H B MERRILL

Emulsions and suspensions in the leather industry. A A CLAPIN. *J Am Leather Chem Assoc* 25, 544-50(1930)—An address. H B MERRILL

Photomicrographic methods for determining the kind of leather and its defects. F FEIN. *Collegium* 1930, 503-8.—The method is briefly described. Eight photomicrographs ($\times 25$) of sheep, horse, calf and goat leathers are reproduced. I D C.

The international commission on the analysis of artificial bates. V. KUDPLKA. *Collegium* 1931, 16-21.—The literature since 1928 is reviewed and the problems yet to be solved are stated. A number of substances have been proposed as substrates but casein is most used. The casein method is not yet satisfactory. Also in *J Intern Soc Leather Trades Chem* 15, 56-60(1931). I D CLARKE

Enumeration (counting) of bacteria in soak waters. GEO F ROCKWELL. *J Am Leather Chem Assoc* 26, 2-7(1931)—In sampling, a piece of skin should be included. Correct dish may vary from 1 to 1,000,000, depending upon the type of soak water. In plating, 1 cc of a fresh meat ext is added to the dish, followed successively by 1 cc dil soak water and the agar. Incubation at 37° for 24 hrs, followed by 24 hrs at room temp, is advised. H B MERRILL

The problem of improving the quality of the hides. A GANSER. *Boll ufficiale staz sper ind pell mout concanti* 8, 110-5(1930). G SCHWÖCH

Contribution on hide damages by M. Bergmann and collaborators. IX. Bacteriology of the reddening of salted hides. WILLI HAUSSAM. *Collegium* 1931, 12-6, cf C A 24, 262.—A description is given of bacteria, other than those described previously, which have been isolated from red stained hides. I D CLARKE

Raw-skin defects and their effect on leather. FRITZ STATNER. *Collegium* 1931, 3-12.—A summary of present knowledge of skin defects and damages. I D C.

Use of sulfide-lime liquors in the sterilization of hides infected with anthrax. M E ROBERTSON. *Leather World* 22, 407, 488, 500, 670, 792(1930). *J Am Leather Chem Assoc* 26, 120, cf C A. 24, 4653 and following abstr. H B MERRILL

Use of sulfide-lime liquors in the sterilization of dried hides infected with anthrax.

IV. Experiments on China hides. DOAGTHV JORDAN LLOYD AND MADGE E ROBERTSON. *J Intern Soc Leather Trades Chem* 14, 641-57(1930), cf C A 24, 4653.—Pieces of hide known to be infected (from guinea pig inoculations) were treated with lime sulfide soln at various temps for different time periods, neutralized triturated, extd, the ext was centrifuged, and guinea pigs were inoculated with the deposit. The proportion of positive results increases with the vol of ext taken for centrifuging showing that a relatively small no of organisms are very resistant to treatment. Complete sterilization, with little or no damage to hides, was attained by treating them for 8 hrs at 32° in a soln contg 10% Na_2SO_4 + 1.2% CaO followed by a 4 day treatment at room temp. Higher temps and (or) sulfide concns ruined the hides. Fairly certain disinfection results at temps down to 20° if the time is increased, but below 20° disinfection is very slow and uncertain. Cf preceding abstr. H B MERRILL

Estimation of total sulfuric acid in hides. X. K. A BUSHMAN. *Ukrainskii Khim Zhur* 5, Tech Pt. 113-5(German Abstract 116)(1930)—The method is based upon the chlorination principle applied by A P Semenztev and V L Pavlov to the detection of As in the intestinal tract of animals and man (*Arch Crim and Med Jurispr.* 1, 43(1928)). Place 4-5 g of finely cut hide in a 15-1 flask, add 25 cc distd water and 10 g NaOH, heat gradually to boiling and when the hide is almost dissolved add 80 cc water and continue boiling till the hide is completely dissolved. Cool in ice water and introduce a moderate stream of Cl_2 gas for about 15 min. From time to time, add distd water until a final vol of about 150 cc is reached to prevent sepn of NaCl crystals. At the end of 15 min, acidify with HCl and remove the Cl_2 gas by bubbling air through the soln. Filter and to the filtrate add a 3% soln of hot $BaCl_2$. B S L.

Hydration of goat skin. FOWNE R. THURTS AND ADOLF GRINKRAUT *J. Am. Leather Chem. Assoc.* 26, 171-47 (1931).—Hydration during soaking and liming is followed by the dilatometer method (C. A. 23, 2119). With H_2O max hydration is attained in 14 hrs., followed by degenerative dehydration ascribed to bacterial or enzyme action. $HCOOH$ in soak decreases and delays attainment of max hydration, this effect increases with increasing $HCOOH$ concn. the onset of degenerative changes is correspondingly postponed but not otherwise affected. $NaOH$ in soak delays the attainment of, but does not decrease the extent of max hydration and does inhibit degenerative dehydration. Extent of max hydration and time required for its attainment are both decreased by increasing the temp. of the soak. Hydrations produced in soaking and in subsequent liming are largely additive. In soaking dried skins hydration seems to take place in 2 stages the second of which is shorter and results in more hydration than the first.

H. B. MERRILL

Principles of the liming process. R. H. MARRIOTT *Leather World* 22, 305 (1930). *J. Am. Leather Chem. Assoc.* 26, 110. The liming process has 2 functions: to remove the hair and to plump and set the fibers in prep. for tanning. The conditions of liming should be regulated so that both aims are achieved simultaneously. Plumping of the fibers results in a characteristic "wave" which is described as flat or steep; the optimum condition depends upon the kind of leather. Puhuring results from the action of alkali on keratin which is promoted by reduction of cystine linkages before the alkali has acted.

H. B. MERRILL

[Some] details of leather manufacture. IRENE A. FINNA *Hide and Leather* 80, No. 11, 21 (1930). Soaking green salted skins prior to liming is unnecessary. $NaCl$, blood, etc., carried into lime when soaking is omitted as without detrimental effect. Omission of soaking avoids bacterial damage. For dried hides, neither acid soaking nor alk. soaking is to be recommended. Salts are effective in the order $NH_4Cl > (NH_4)_2SO_4 > CaCl_2 > NaCl$. Some tanners put dried hides directly into strong Na_2S and lime, for combined soaking unbaring with good results. A soln. contg. 1.25% NH_4Cl and 1.25% Na_2S (60%) is an effective softener.

H. B. MERRILL

Processes of patent-leather manufacture. IRENE A. FINNA *Leather World* 22, 36, 110, 100, 412, 593 (1930). *J. Am. Leather Chem. Assoc.* 26, 118. A description.

H. B. MERRILL

Hazards in the manufacture of patent leather. H. B. MERRILL *Hide & Leather* 80, No. 14, 19 (1930); *J. Am. Leather Chem. Assoc.* 26, 3d. The naphtha fire hazard is discussed.

H. B. MERRILL

Acidity of leather. H. G. BENNETT *J. Intern. Soc. Leather Trades Chem.* 15, 31-8 (1931).—The "acid figure" of Atkin and Thompson (C. A. 23, 4513), which is believed to be the pH value of the concd. soln. of tannin on the outer part of the fibers, can be obtained arithmetically as well as graphically from measurements of pH value at known dilns. in 0.1 N KCl . pH (at any diln.) = $\tan \theta \log d + F$, where $\tan \theta$ = slope of the pH diln. plot, d = diln. and F = acid figure, or pH at zero diln. $\tan \theta$ is thought to be fully as important as F . Two points are sufficient to det. the pH diln. plot, and calens. are simplified if these 2 pH values whose logs are simple integers are detd. at fixed dilns., e. g., diln. = 100 fold and 1000 fold. $\log d = 2$ and 3. To find the proper wt. of leather to take to obtain these dilns. when the leather is soaked in 100 cc. of 0.1 N KCl , det. the % H_2O (H) and weigh 101 H' g. for 100-fold, and 101 H' g. for 1000 fold diln. If pH values of these solns. are $(pH)_1$ and $(pH)_2$ then $F = 3(pH)_1 - 2(pH)_2$, and pH (at any diln.) = $[(pH)_1 - (pH)_2] \log d - F$. Examn. of many leathers (chiefly sole) selected as representative com. products showed only 1 leather for which $F < 2.5$, and av. value of F was 3.5. Generally a high value for F is assocd. with a low value for $\tan \theta$; that is, with a well buffered leather and this is due often to the presence of appreciable amts. of Na salts.

H. B. MERRILL

Application of the Atkin-Thompson method for acidity to vegetable-tanned calf leather. MENAY B. MERRILL AND RUDOLF G. HENRICH *J. Am. Leather Chem. Assoc.* 25, 511-7 (1930), cf. C. A. 23, 4543.—Data are presented which indicate that the method gives results on leathers tanned with so-called spruce ext. (which is high in sulfonated bodies) that are rational and self-consistent, in contrast to results given by the Procter-Searle method.

H. B. MERRILL

Present state of the determination of wear resistance of leather and methods for bettering it. U. J. THURAU *J. Intern. Soc. Leather Trades Chem.* 14, 602-23 (1930); cf. C. A. 23, 1768, 4090; 24, 2931, 4654.—An improved wear test machine is described. For standardization, a plate of pure Cu , previously heated to 600° for 15 min., is employed, and all wear resistances are expressed in terms of this standard taken as 100. Soles are rubbed across a Carborundum plate under pressure of 70 kg. (av. human wt.),

Accessory gadgets assure a clean and uniform wearing surface at all times. Results are expressed in cm^2 of abrasive surface with which 1 cm^2 of leather must be brought into contact in order to wear off a layer 1 mm. thick. As the ratio of abrasive surface to leather surface exposed at a single revolution is 1:1, the coeff. of resistance is directly proportional to the no. of revolutions and inversely proportional to the thickness removed. The latter is obtained from wt. loss and density measurements. The coeff. of resistance varies with time, generally increasing after the grain layer has been worn away, curves are given for numerous leathers. Results are given for numerous kinds of sole leather and substitutes. Wear resistance of vegetable tanned leather is increased by retannage with chrome, by previous acid swelling of the pelt, and by increasing its resistance to extn. by H_2O . The superiority of leathers tanned by the "slow" process is ascribed to their relatively low content in matter extractable by H_2O . Hardness is measured by the Brinell machine. I. B. MERRILL

Little trips to tanneryland. L. Wild grain. JOHN A. WILSON. *Hide and Leather* 81, No. 7, 8(1931).—Wild grain is caused by overdevelopment of the blood vessel system supplying the sebaceous glands. It can be minimized by proper control of all tannery operations. I. B. MERRILL

Report of the commission on qualitative analysis of tannins. O. GERNIGROSS. *Collegium* 1930, 521-32; cf. *C. A.* 25, 230.—The aniline- HCl , cinchonine and fluorescence tests for sulfite cellulose were generally in agreement. Pine and malt to bark exts. interfere with the fluorescence test in its present form. I. D. CLARKE

Report of the commission on quantitative tannin analysis. M. BERGMANN. *Collegium* 1930, 573-4; cf. *C. A.* 23, 4592.—A review. I. D. CLARKE

Rapid tanning. MAX BERGMANN, WILLY MONZ AND LUDWIG SELIGSBERGER. *Collegium* 1930, 530-4.—Neradol N D and α and β naphthalenesulfonic acid solns. pass through skin more rapidly than H_2O does, therefore they enlarge the skin capillaries. But the total H_2O of swelling and imbibition is less and so the H_2O of swelling must be less (H_2O of imbibition is greater since the capillaries are larger) in these solns. than in H_2O . Agents to increase the speed of tanning may not only decrease the size of the tannin particles but may act on the hide to decrease resistance to penetration. Also in *J. Intern. Soc. Leather Trades Chem.* 15, 67-72(1931). I. D. CLARKE

The enzymes of tanning chemistry and the structural-chemical explanation of their action. MAX BERGMANN. *Collegium* 1930, 516-20.—Studies of some reactions of protein building units are reviewed, these studies should aid in explaining the mechanism of those reactions of proteins which are catalyzed by enzymes. I. D. CLARKE

Filtration of tannin solutions for the determination of insolubles. A. L. C. A. Committee Report, 1930. J. S. ROGERS, et al. *J. Am. Leather Chem. Assoc.* 25, 493-510(1930).—For the purpose of this investigation, "insolubles" are defined as substances which render the tannin soln. of analytical strength, at 18° , cloudy, turbid or opalescent by either reflected or transmitted light. Methods studied are (1) the International Method and the methods of (2) Turnbull (*C. A.* 23, 2596), (3) Riess (*C. A.* 23, 3369, 4592) and (4) McCandlish Atkin (*C. A.* 23, 3368); solid quebracho, chestnut, hemlock and "valex" exts. were used. In all cases the solns. were filtered until a clear filtrate was obtained, after which 2 successive portions were collected, pupetted, evapd. and weighed. Results (regarded as preliminary) indicate that the International Method returns too low insolubles and does not always give clear filtrates, that the other 3 methods do give clear filtrates, and that of these the Riess method is most promising from the standpoint of app. and manipulation. I. B. MERRILL

The tannin balance in the extraction of pine bark. E. BELAYSKY AND G. WANKE. *Ledertech. Rundschau* 23, 1-4(1931).—The tannin balance for pine bark and the ext. showed a gain of 8.38% if the raw and spent bark were analyzed by the International Method but only 0.39% if the method was modified to give max. results. Completeness of extn. is affected by many factors such as temp., time and the percentage which the ext. is of the raw material but the greatest loss of tannin is always in the extd. bark. I. D. CLARKE

Determination of the copper content of tanning extracts. Committee report. D. BURTON. *J. Intern. Soc. Leather Trades Chem.* 15, 99-106(1931).—Cu in exts. is even more objectionable than Fe. Cu is always extd. in the ext. ash. Of numerous procedures tested, electrolytic deposition from H_2SO_4 soln. is recommended. Procedure.—Dry 50 g. ext. in an oven, ash at min. temp., moisten with H_2SO_4 , expel chlorides, dissolve in 10 cc. 2 N H_2SO_4 and dil. to 100 cc. Electrolyze overnight at 2 v. Wash with H_2O , then with EtOH (abs.), dry at 100° and weigh. I. B. MERRILL

The determination of the color of tanning extracts. G. BALDRACCO. *Boll. ufficiale staz. sper. ind. pell. mat. concianti* 8, 103-9 (1930).—The different methods are reviewed. The Pulfrich stage photometer is considered as the instrument most suited for the detn. of the color of tanning exts. G. SCHWOCH

Determination of the p_n (value) of tanning extracts. D. BURTON, et al. *J. Intern. Soc. Leather Trades Chem.* 14, 567-601 (1930).—The quinhydrone method is unsuitable. The bubbling H_2 method gives accurate results with exts. that are not sulfited. The glass electrode method (C. A. 23, 4365) is suitable for use in all types of tannery liquors. Working directions for the 2 latter methods are given. H. B. MERRILL

Importance of the p_n of the pelt in tanning. G. PLACK. *J. Intern. Soc. Leather Trades Chem.* 15, 22-4 (1931).—It is a mistake to assume that the pelt speedily assumes the p_n value of the tan liquor in which it is placed, and in center attention on the p_n value of the liquor. Heavy hides, especially, should be brought to the optimum p_n value (which generally lies in the p_n range 5 to 3) before entering the tan liquors. This procedure makes it unnecessary to add acid to the tan liquors, permits rapid tannage and satisfactory plumping in liquors of moderate acidity, greatly simplifies control and makes for numerous operating economies. H. B. MERRILL

Determination and control of the buffer index of tan liquors. WINIFRED B. PLEASS. *J. Intern. Soc. Leather Trades Chem.* 15, 73-8. *Leather World* 23, 35 (1931).—The detn. of effective acidity (p_n) should be accompanied by that of reserve acidity (buffer index). The latter is defined as "the no. of cc. of N NaOH or HCl required to be added to 100 cc. of a tan liquor of 20° barkometer (sp. gr. = 1.020) to alter its p_n value by 1 unit". The buffer index should be estd. on a portion of the titration curve covering at least 2 p_n units, in the p_n range 2.5-6. This involves titration with NaOH for liquors of low p_n , and with HCl for liquors of high p_n . Liquors of p_n value = 3.0-4.5 must be titrated with both. If the barkometer of the undiluted liquor is less than 20°, the buffer index at 20° is obtained by simple proportion. Generally speaking, a high buffer index is given by old, mellow liquors, rich in non tannin or sol. Ca salts (such as bottom suspenders). The buffer index should decrease from the oldest liquor in a series to about the mid point, and should be approx. const. for the remaining (fresher) liquors. Sharp differences in buffer index between successive liquors are objectionable. The buffer index increases with increasing non tannin/tannin ratio in the tanning materials used, is increased by adding lactic acid, and decreased by adding H_2SO_4 . A high buffer index results in the production of a soft, full leather; a low buffer index results in a firmer, harder leather. H. B. MERRILL

Determination of p_n value of tannery liquors. W. B. PLEASS. *Shoe & Leather Reporter* 80, No. 12, 21 (1930). *J. Am. Leather Chem. Assoc.* 26, 111.—The glass electrode is recommended as most reliable, with the capillator second. H. B. MERRILL

The basicity concept of one-bath chrome liquors. R. STIASNY. *Collegium* 1930, 574-7.—The basicity of the Cr salt in a one-bath Cr liquor is greater than that of the liquor as a whole because of "olition", the complex ol. compds. react with the free acid only very slowly. It is important to know the basicity both of the liquor and of the Cr salt. The former is given by the formula $[(a-b)/b]100$ and the latter by $[(a-b+c)/a]100$, in which a is the no. of cc. of 0.1 N $Na_2S_2O_8$ used in the iodometric detn. of Cr (in 25 cc. of liquor), b is the no. of cc. of 0.1 N NaOH used in titrating the liquor as usual (to phenolphthalein) and c is the no. of cc. of 0.1 N NaOH used in a cold titration to p_n 3.0 with p -benzenesulfonic acid n -benzylamide. At p_n 3 only the free acid in soln. is detd. I. D. CLARK

Origin of ammonias found in lime liquors. R. H. MARRIOTT. *J. Intern. Soc. Leather Trades Chem.* 15, 25-30 (1931).—Collagen from oxide and degraded oxhair was treated with satd. $Ca(OH)_2$ for periods of 8-72 days. Total N and volatile N were detd. in the used liquors. The undissolved proteins were washed and hydrolyzed by 20% HCl . Hydrolyzate was analyzed for NH_3 and total N. With collagen, the amt. of NH_3 found in the liquor agrees almost quantitatively with the observed decrease in amide N in the protein, with hair, more NH_3 is formed than can be accounted for by the decrease in amide groups, the excess is supposedly due to breakdown of cystine groups. Amide N is removed very much faster both from hair and from collagen than is non amide N. H. B. MERRILL

Preparation and preservation of sheepskin skivers for color tests [on tanning materials]. Committee Report, 1929-30. C. A. BLAIR, et al. *J. Am. Leather Chem. Assoc.* 25, 550-61 (1930).—Two methods for depickling and preservation are investigated with pickled skin as the starting material. (1) The skin is suspended in 5% $NaCl$ soln. and the acid titrated repeatedly to the neutral point, until the skin no longer swells when placed in H_2O . The skin is preserved in satd. $NaCl$ soln. (2) Excess

borax is used and the excess neutralized with dil AcOH. The pH is adjusted to 4.8 with acetate buffer. The skin is preserved in phenol boric acid soln. pH values of solns in equal with skivers prep'd according to (1) are not very uniform. Method 2 is generally preferred, although the borophenol causes some darkening of the skiver on long storage. H. B. MERRILL.

Treatment of tannery wastes. Condensed report to the Sanitary Water Board. TANNERY WASTES DISPOSAL COMMITTEE OF PENNSYLVANIA. *J Am Leather Chem Assoc* 26, 70 (1931).—Two main plans of waste treatment were developed and given long tests under full scale operating conditions in a small vegetable leather tannery. Depending upon the degree of purification required, which in turn depends upon the type of wastes, type of stream and vol. of wastes compared to stream vol., 1, 2, 3 or all 4 of the following steps may be required: mixing and settling, primary filtering and settling, chem. coagulation, secondary filtering and settling. The above is a general scheme believed to be applicable to any tannery. A special scheme consists in segregating and chemically pptg. certain wastes, followed by filtering and settling. Spent tan liquors and bleaches are the forms of pollution most difficult to treat, and arrangements should be made to discharge them at a uniform day-to-day rate. Sludge should be removed at regular intervals, cinder bed drying is satisfactory. Cost data and detailed drawings of the exptl. plants are given. H. B. MERRILL.

Tanning shark skins. F. A. COOVERS. *Hide & Leather* 80, No. 11, 24 (1930), *J Am Leather Chem Assoc* 26, 110.—A description. H. B. MERRILL.

Nature of water solubles in leather tanned with wattle bark extract. H. R. O. PAGE AND H. C. HOLLAND. *J Am Leather Chem Assoc* 26, 143-56 (1931), cf. *C A* 23, 541.—Soaked, limed and bated steer hide was tanned for 6 weeks with wattle bark ext. at pH 4; thereafter it was tanned at pH values of 4, 3 and 2 for periods up to 2 yrs. Analyses were made at the end of 3, 6, 26, 52 and 104 weeks, the authors' method for free and combined H_2O -sol matter being used. At the outset, fixed tannin rises rapidly to 30 g. per 100 g. collagen, with little combination of H_2O -sol matter, in the 2nd stage, tannin slowly increases and combined H_2O -sol matter rises to about 30 g. per 100 collagen, in the final stage (3 months to 2 yr.) fixed tannin continues to rise, approaching a total of 69 parts per 100 collagen, while combined H_2O -sol matter remains stationary. The pH value in the 2nd and final stages is without effect on combined H_2O -sol matter, fixed tannin increases slightly with decreasing pH value in the range studied. In expts with hide powder, with different ratios of wattle ext. solids to collagen, the ratio, fixed tannin: combined H_2O -sol matter, decreases steadily with increasing ratio of ext. to collagen. The percentage of the total ext. solids removed from the solo is nearly independent of the ratio of the ext. to collagen, decreasing slightly in the more concd solns, which show the presence of tannin by the gelatin salt test after tanning is complete, for the more dil solns, the % of the total ext. solids thus removed as fixed tannin and combined H_2O -sol matter is very nearly the % of tannin in the ext. according to the A. L. C. A. Method. These results are explainable on the assumption that collagen combines with tannin to form a compd. highly resistant to washing (fixed tannin), then with more tannin to form a compd. less resistant to washing (combined H_2O -sol matter). The fixed tannin and combined H_2O -sol matter are thus the same substances differently combined. The A. L. C. A. Method for detg. tannin is approx. correct, that of Wilson and Kern is in error. The effect of the pH value was studied at 15° and 35° for 6-weeks' tannage. At 15° both fixed tannin and combined H_2O -sol matter are at a min at pH = 5. At 35° this min. vanishes, and fixation decreases from pH = 3 to pH = 8. Fixed tannin is much greater, and combined H_2O -sol matter somewhat less at 35° than at 15°. By treatment with dil H_2SO_4 at temps. from 35° to 60°, most of the combined H_2O -sol matter is fixed as tannin. When a leather so treated is retanned, fixed tannin does not increase but combined H_2O -sol matter reassumes its former value (before the heat-acid treatment). The change of state of combination is ascribed to changes in the collagen and not in the material fixed originally as combined H_2O -sol matter. The recurrence of the factor 30 in the ratios of fixed tannin and (or) combined H_2O -sol matter to collagen suggests that the combining ratio of wattle tannin to collagen is 100:30, giving a combining wt. of 225 (if collagen is 750) for wattle tannin. Conclusion: Combined H_2O -sol matter is tannin, combined in a manner different from fixed tannin. H. B. MERRILL.

Deterioration of vegetable-tanned leather in storage. R. PARADAY JONES. *J Intern Soc Leather Trades Chem* 14, 624-40 (1930).—Old book bindings were analyzed for H_2SO_4 (Procter Searle method), "difference figure" (*C A* 23, 4367), sol. N and grease. Comparison of backs (exposed) and sides (less exposed) showed that large amts. of SO_2 are absorbed from the atmosphere, particularly in industrial cities. Ex-

amination of books kept side by side for many years shows that leathers differ in their absorptive capacity for SO_2 , the absorption may be a result rather than a cause of the rotting. All rotted bindings were very high in H_2SO_4 , but not all leathers high in H_2SO_4 were rotted. An extreme case is a leather contg 4.7% H_2SO_4 , which was sound after 42 yr. Sol. N is proportional to H_2SO_4 , and like it is not directly related to deterioration. It is thought that the kind of tanning material detrs whether or not a large H_2SO_4 content is harmful. Other factors may be phys condition of fibers resulting from pre-tannard operations, lack of grease, overtanning. It is advised that bookbinding leathers be tanned with oak bark or sumac only, and that H_2SO_4 be not used after tanning is complete. The use of H_2SO_4 for deliming and pickling has no ill effect.

H. B. MERRILL

The action of pickles of different compositions. II. Acid absorption and swelling of collagen in sulfuric acid and formic acid pickles. A. KONTZEL AND W. PREISSENTANZ. *Collegium* 1930, 577-83. cf. C. A. 24, 4654. A H_2SO_4 pickle, whether it contains NaCl or Na_2SO_4 , behaves like a HCl pickle. The absorption values for H_2SO_4 are the same as those for HCl , the calcs given previously for HCl apply to absorbed H_2SO_4 in H_2SO_4 , Na_2SO_4 pickles without reservation but for H_2SO_4 , NaCl pickles the NaCl influences the absorption of H_2SO_4 , so that a correction (max. absorption is 0.6 instead of 0.7 equivs.) must be applied. The swelling weights of collagen are analogous to those in HCl . NaCl pickles the abs. values, however become smaller as Cl^- is replaced by SO_4^{--} (swelling in HCl NaCl > in H_2SO_4 , NaCl > in H_2SO_4 , Na_2SO_4 pickles). Systems contg weak org. acids are entirely different. The rising portion of the absorption curve is not linear and the absorption of acid is either greatly decreased (lactic) or greatly increased (acetic, formic) by the salt content of the pickle. The pickling or de swelling actions of NaCl and Na formate on formic acid are exactly the same although the former has no, and the latter a great, influence on the pH . These results disprove Procter's theory that all pickles contg NaCl , regardless of the kind of acid, behave like HCl pickles. They also disprove the theory that the swelling action of weak org. acids can be repressed by buffering. Salts repress swelling because of their pickling action and regardless of whether they have a buffering action or not.

I. D. CLARKE

Report of the oil and fat commission. M. AUERBACH. *Collegium* 1930, 532-4. — A brief discussion of general methods.

I. D. CLARKE

Dyestuffs and leather. W. F. NORTON. *Leather Trades Rev.* 63, 640, 739 (1930), *J. Am. Leather Chem. Assoc.* 26, 51. — Uneven coloring on Cr leather is ascribed (1) to uneven acid distribution in leather due to prolonged "horning" before coloring and (2) to the presence in the dye mixt. of dyes that penetrate and fix at different rates. Over-mordanting increases rubbing of the color. Sufficient penetration must be secured so that the leather does not appear lighter in shade when it is stretched. Penetration is secured by using acid dyes, adding acid to the dye bath and slow drying of the dyed leather. Basic and direct dyes should be used for topping off—not for bottoming, and they should never be used in a bath with acid dyes. Practical working directions are given for various types of leather and special color effects.

H. B. M.

Osage orange and yellow wood extract. G. DESMURS. *Leather World* 22, 201 (1930), *J. Am. Leather Chem. Assoc.* 26, 116. — Formulas are given for using yellow or Cuba wood ext. in dyeing Cr leather. The following test is used for differentiating osage orange from wood. Measure 10 cc. of 0.4% ext. into a test tube, and add 20 cc. of a 10% U nitrate soln. Colors produced are: osage orange, a full orange-red; Cuba wood, yellow-orange. Other tests are given.

H. B. MERRILL

Technical preparation of glue and gelatin. D. GURIAN. *Kolloid-Z.* 54, 239-43 (1931); cf. C. A. 25, 1408. — A review of German patents.

A. FLEISCHER

Influence of pancreatin on collagen in the absence of neutral salts and buffer mixtures (KONTZEL, DIETSCH) 11A. Tests on artificial leather (HESSE) 26. Glass electrode and vacuum-tube potentiometer [in tannery analyses] (CAMERON) 1. Treating fibrous materials with rubber, etc. [for making artificial leathers] (Brit. pat. 337,359) 25. Protein products resembling horn (Brit. pat. 338,015) 25. Apparatus for continuous filtration and dehydration of pulp for artificial leather manufacture (Brit. pat. 337,346) 23.

Dressing leather. OTTO L. STEVEN. Ger. 514,723, July 23, 1927. See U. S. 1,750,732 (C. A. 24, 2635).

Leather. HERMANN BOLLMANN and BRUNO REWALD. Ger. 517,353, Dec. 8, 1927, and 517,354, Feb. 24, 1928. A substitute for yolk of egg in the manuf. of glacé leather comprises an aq. soln. of an aromatic or aliphatic sulfonic acid or sulfonate and a phos-

phatide of vegetable or animal origin with or without a vegetable albumin, e. g., soybean albumin. An example is given (517,353). Alternatively there may be used an emulsion of vegetable or animal albumins and phosphatides prep'd by dissolving the albumins in water contg a small quantity of alkali and then adding the phosphatides, which may be first dissolved in a fatty oil (517,354).

Leather. CARL H. TRÖGER, JR. Ger. 517,102, July 26, 1928. Vegetable tanned leather is treated before drying, with a bleaching bath to which substances promoting fermentation e. g., yeast, have been added. In this way the time required for drying is considerably reduced.

Tanning. W. SAILER. Brit. 337,377, April 30, 1929. Hides are treated with a liquor prep'd by treating sulfite cellulose lyes with oxidizing compds. such as perborates or persulfates followed by treatment, about 2 hrs. later, with Cl or Br. Pelts may be preliminarily treated with the liquor and then depilated. Cf. C. A. 25, 2019.

Rotary-drum apparatus for tanning skins. F. G. WILSON. Brit. 337,408, July 22, 1929. Structural features.

Tanning agents. J. G. FARBERVIND A. G. (Ludwig Teichmann and Hermann Noerri inventors). Ger. 517,446, Jan. 3, 1929. See Fr. 687,411 (C. A. 25, 839).

Tanning agents. J. R. GERV A-G. Swiss. 141,788 and 141,789, July 30, 1928. Adds to 138,854 (C. A. 25, 839). A weakly acid powerful gelatin depositing tanning agent is obtained by the reaction on β -chloromethylbenzylsulfonic acid in dil. AcOH suspension with Na naphthalene β -sulfonate or Na *o*-cresol β -sulfonate. Cf. C. A. 25, 839.

Tanning hides. O. RÖHM. Brit. 337,524, Dec. 6, 1928. A substitute for egg yolk suitable for use in the manuf. of glaze leather consists of an oil emulsion and one or more esters formed from di- or polyhydric alcs. and H_2PO_4 . These constituents may be used in a tanning liquor together with meal, alum and NaCl.

Tanning skins. J. R. GERV A-G. Ger. 514,674, July 7, 1925. The depilated skins are treated with mineral acid and neutral or basic Cr, Al or Fe salts. They are then treated, in the same bath or in a separate soln., with neutral synthetic tanning material, the soly of which, in water, depends on the presence of the H_2SO_4 group. Sulfite cellulose lye may be also used for further steeping of the skins. In an example, the skin is pretreated with Cr alum and tanned with the condensation product of CH_3O and cresolsulfonic acid. Other examples are given.

Moistening and "wetting out" leather and hides. FRIEDRICH POSPIECH (to Chemische Fabrik Pott & Co.). U. S. 1,794,920, March 3. Thorough moistening is effected with an aq. soln. contg sulfonic acids of a compd. obtained by condensing naphthalene or a ring substitution product of naphthalene such as naphthalenesulfonic acids in the presence of condensing and oxidizing agents such as H_2SO_4 with an alc. of the aliphatic series contg. over 2 C atoms such as Pr or Bu alcs., cyclohexanol, etc.

Glue. BENNO NORDON. Ger. 517,098, Dec. 30, 1928. Glue liquid in the cold is prep'd by treating animal glue with $HClO_4$. Thus animal glue 10 dissolved as usual in water 15 may be stirred with 20% $HClO_4$, 3 parts.

Glue and gelatin. HERMANN STREIDL. Ger. 514,721, Feb. 7, 1928. App. for molding is described.

30—RUBBER AND ALLIED SUBSTANCES

C. C. DAVIS

The preparation of plantation rubber. G. MARTIN. *Bull. Imp. Inst.* 28, 440-50 (1930).—A description is given of the methods of prep'n. of plantation rubber used in Ceylon, Malaya and Java, with tables to show the differences in detail between the procedures followed in the 3 countries. Also in *Trop. Agr.* (Ceylon) 76, 18-26 (1931).

Liquid rubber compounding. WEBSTER NORRIS. *India Rubber World* 83, No. 5, 53-5, 64 (1931).—An illustrated description of recent developments in the preservation, diln., stabilization and neutralization of latex, mixing operations, typical compds. and industrial applications.

Action of nitrogen oxide on rubber. ADOLF GORGAS. *Ber.* 63B, 2700-5 (1930).—Various observers have reported the formation of definite products by the action of NO_2 on rubber (Weber, $C_{11}H_{16}N_2O_4$, Emden, $C_{12}H_{16}N_2O_4$, Harnes, Alexander, products of more complex compn.). These workers paid no particular attention to the purity of the NO_2 . The best results were obtained with the gaseous decompn. prod-

acts of $\text{Pb}(\text{NO}_3)_2$. Imlen used the product obtained from HNO_3 and starch. G. on the other hand, carefully purified his NO_2 , which was prepd from HCHO and concd. HNO_3 . The resulting mixt of N_2O_4 and N_2O was liquefied in a freezing mixt., completely converted into N_2O_4 with air, frozen repeatedly, crysd and fractionated by distn. For the reaction with rubber the N_2O_4 was used in CCl_4 soln, satd at room temp. Total rubber, prepd from "Revertex" according to Pummerer and Pahl by treatment with alkali at 50° and subsequent dialysis was used in 1% soln in CCl_4 . This rubber, as has been described elsewhere (C. A. 23, 1525), gave by the modified Huns method I nos. agreeing well with the calcd. values and the sol and gel rubbers prepd from this total rubber also gave normal I values (1 mol I per C_{15}H_2 group). Contrary to Pummerer and Mann G. found for the gel rubber with Brl an I no. about 10% too low instead of too high. By the Kaufmann (SCN)₂ method, the 2 rubbers showed no appreciable difference. The nitration was effected by adding the rubber soln to 5 parts of the N_2O_4 soln at 0° and letting the mixt stand 15 min in 1 series of expts and 21 hrs in another. The compn. of the resulting products did not correspond to any of the formulas given by earlier workers, even after pptn. from Me_2CO with Et_2O , and the expts. showed decisively that abs. pure NO_2 does not form homogeneous products with purified rubber. As already pointed out by Alexander the aiddn of NO_2 is accompanied to a considerable extent by oxidation. The aiddn reaction apparently does not give NO_2 derivs. at least, G. has thus far been unable to effect a reduction to NH_2 derivs., but he did observe that when the product of the 21 hr. treatment with NO_2 is heated with concd. alkali about $1/2$ of the N is split off as NH_3 , which makes it probable that the N is present as an NOH group (probably as an isomeric ketone). In view of the strongly oxidizing action of NO_2 expts. were made with NO . As it reacts much more slowly than NO_2 , the well purified NO was passed for 4 hrs. through a very dil. soln. of rubber in CCl_4 at 20° . The product was non pulverizable, and had a compn. (about $\text{C}_{15}\text{H}_2\text{N}_2\text{O}_2$), indicating that oxidation to a marked degree accompanied the addn. reaction. The product was also no more homogeneous or const. in compn. than that obtained with NO_2 , and gave NH_3 with hot concd. alkali. With HNO_3 it is possible to keep the 2 reactions sep. Dil. HNO_3 and rubber give a N contg. product which, however, is partially insol. in Me_2CO , and this insol. portion is N free while the N in the sol. portion is split off as NH_3 by boiling alkalis. The formation of the N contg. product can be prevented by complete removal with urea of the HNO_3 which is formed. The yellow pulverizable product is sol. in PhNO_2 , PhNH_2 , and dil. alkalis. G. hopes soon to be able to report its compn., mol wt., etc. C. A. R.

The rubber era. T. J. GARNER. *India Rubber J.* 80, 643-8 (1930).—An illustrated description is given of modern technical applications of rubber. C. C. DAVIS

The trend of American rubber development research. ARNOLD S. SMITH. *India Rubber J.* 80, 664-7 (1931).—An English version (slightly abbreviated) of C. A. 24, 4954. C. C. DAVIS

Industrial uses of rubber. H. F. FRITZ. *Paper Mill* 54, No. 3, 14, 10-7 (1931).—An outline of the widening applications of rubber rendered possible by the discovery of the Vulcalock process of bonding rubber to metal, which permits of utilizing to better advantage the corrosion resisting and abrasion resisting qualities of rubber. A. P.-C.

Preparation of sheet rubber. R. O. BRISOR. *Malayan Agr. J.* 19, 14-21 (1931).—A discussion is given of present-day methods for the collection and handling of latex, its coagulation, handling of the coagulated rubber and drying and packing of the rubber. C. C. DAVIS

Toxic substances in the rubber industry. XIX. Sulfur and its derivatives. P. A. DAVIS. *Rubber Age (N. Y.)* 28, 513-4 (1931), cf. C. A. 25, 1117.—With proper ventilation, choice of individuals and cleanliness, no trouble is likely with S. Fifteen yrs' experience showed no serious cases of conjunctivitis, dermatitis, etc., from exposure to S. C. C. DAVIS

Technical note on the use of certain solvents in solutions. H. COULANGLON. *Caoutchouc & gutta-percha* 27, 15274-6 (1930).—When a film of cement dries, the evapn. of the solvent causes a lowering of the temp., which may be great enough to reach the dew point and thus cause moisture deposit on the rubber. This results in poor adhesion when the surfaces are united. Since the lowering of the temp. varies with the solvent, it is possible to avoid this condensation of moisture by the proper choice of solvent for the conditions. Data were obtained on the lowering of temp. (wet bulb) of cements of the same viscosity made with solvent naphtha, toluene, aviation gasoline and benzene, resp., with the atm. at different temps. ($10-30^\circ$) and humidities. By plotting the results graphically, the resulting chart shows under what conditions of temp. and humidity cements made with the various solvents can be used successfully without re-

course to warming, and under what conditions and to what extent warming is necessary

C. C. DAVIS

Paragutta, a new insulating material for submarine cables. A. R. KEMP. *J. Franklin Inst.* 211, 37-57(1931).—Gutta percha and balata have proved themselves eminently suitable for insulating long deep sea telegraph cables, but their dielec losses are too high to meet the requirements of submarine telephone cables for long distances or for shorter cables with carrier currents. Accordingly an extensive investigation was made of the causes of elec defects of submarine insulation, including a search for an improved material. One result was the development of a new material called *paragutta*, and the present paper describes its properties and tests for detg its utility. *Paragutta* consists essentially of the purified balata or gutta percha hydrocarbons, with a small proportion of waxes. In developing this mixt, it was necessary first to purify the rubber, especially to remove proteins, in order to increase its elec stability on prolonged immersion in water. Rubber from which water sol substances have been removed absorbs no more water than does gutta percha (cf. Williams and K., *C. A.* 21, 1372, Lowry and Kohman, *C. A.* 21, 1372, Boggs and Blake, *C. A.* 20, 2092), but when immersed in water it fails completely as an insulator. To explain this elec instability, rubber hydrocarbon freed of proteins, resins and other impurities was prepd and tested. The product absorbed very little water, and on peolnged immersion in water it showed practically no change in elec properties. Many expts showed that the only practical method for removing proteins from rubber is to heat the rubber in an autoclave in water alone, which hydrolyzes the proteins to water sol products. Subsequent washing removes nearly all N. Sheet rubber or latex can be used for prepg N-free rubber in this way. Thus ammoniated latex is diluted by addn of 4 vols of water, heated 10 hrs at 150° in an autoclave, cooled, coagulated with AcOH, washed thoroughly and dried. The N content is then under 0.1%. When vulcanized, this deproteinized rubber is superior to ordinary rubber in its elec stability in water. The deproteinized rubber is also more easily plasticized and mixed with gutta percha than is ordinary rubber. *Paragutta* is prepd by blending deproteinized rubber with purified gutta percha hydrocarbon, with the optional addn of special hydrocarbons or montan wax. The proportions of the 3 components may be varied to suit the conditions. Thermoplastic insulators thus prepd are almost like gutta percha in mech and elec properties in water, and they are cheaper. Data show tensile, compression, flexibility, plasticity, brittleness, water absorption and elec tests of *paragutta* in comparison with gutta-percha. These data prove conclusively that *paragutta* behaves in almost the same way as gutta percha in all these respects. It has the desirable thermoplastic and mech properties of gutta percha, and its superior insulating properties render it particularly adapted for transoceanic telephone cables. *Paragutta* is also of advantage for shorter deep sea carrier telephone cables and for ocean telegraph lines. C. C. DAVIS

Reconditioning air bags by burning. CHARLES E. MAYNARD. *Chem Met Eng* 38, 91, *Rubber Age* (N. Y.) 28, 565(1931).—A new machine is described for reconditioning air bags used in the vulcanization of rubber tires. It removes the hardened surface layer of used bags by burning and scraping, thus leaving a smooth surface and thoroughly reconditioning in one operation. C. C. DAVIS

Accelerators of vulcanization. Mercaptobenzothiazole. P. JACOB. *Caoutchouc & gutta percha* 28, 15350-4(1931), cf. *C. A.* 25, 1118.—The prepn and properties of mercaptobenzothiazole are described, including its behavior and utility in rubber mixts. Mercaptides. *Ibid.* 15394-6.—By mercaptides are meant combinations of mercaptobenzothiazole with other accelerators having basic properties, e. g., with hexamethylenetetramine, diphenylguanidine, *di-o* tolylguanidine and ethyldiethylaniline. The behavior of some of these com substances as accelerators is described, with examples to show their utility. C. C. DAVIS

Oxidation of vulcanized rubber mixtures extracted with acetone. T. YAMAZAKI AND K. OKUYAMA. *Caoutchouc & gutta percha* 28, 15359-60(1931).—See *C. A.* 24, 4664. C. C. DAVIS

Activated carbons (SIRONIN) 18. Modifying the properties of resins [for rubber substitute manufacture] (Brit. pat. 337,733) 26. Modifying isocolloids for rubber-substitute manufacture (Brit. pat. 337,732) 26. Improving elasticity of knitted artificial silk fabrics [for waist bands, garters, etc.] (Brit. pat. 337,400) 25. Waterproof coatings (Swiss pat. 142,172) 18. Rubber-coated fabric (U. S. pat. 1,795,199) 25. Means for heating or intensively cooling hollow rolls or drums for working rubber (Ger. pat. 517,361) 1. Rubber surfaced paving blocks (U. S. pat. 1,794,220) 20. Solvents for

rubber (Ger pat 517,007) 13. Treating fibrous materials with rubber, etc [for making tire fabrics] (Brit pat 337,350) 25.

Rubber. I G FARREND A G (Julius Eisele and Johannes Stübel, inventors). Ger 517,490, Sept 13, 1927 Rubber latex is rapidly mixed with sufficient acid to establish a hydrium concn greater than $pH 2$, preferably greater than $pH 1$, and a compd. reducing the surface tension is then added, e g, Na β naphthalenesulfonate or Turkey-red oil. The mix sets to an irreversible gel with sepn of a small quantity of clear serum. Vulcanizing agents, accelerators or fillers may be added before or after the formation of the gel. Examples are given. Cf C A 25, 1706.

Rubber. SOCIETÀ ITALIANA PIRELLI. Ger 517,208, Oct 12, 1926 The aging properties of rubber are improved by addn of about 5% of the evapn residue of latex serum from which quebrachitol has been removed, e g, by crystn. Albumins may also be removed from the serum to be evapd, e g, by pptn with tannic acid.

Rubber deposition. IMPERIAL CHEMICAL INDUSTRIES, LTD., and R B P P CLARKE. Brit 317,446, Oct 21, 1929 A hollow gas permeable vessel, filled with CO_2 under pressure, is immersed in latex, and rubber is deposited on it as the CO_2 diffuses through the walls of the vessel and effects coagulation of the rubber deposit. Material to be impregnated may be wrapped around the vessel and the vessel may be shaped to produce sheets or articles of other desired form.

Rubber composition. PILWOOD B SPEAR and ROBERT L MOORE (to Thermatomic Carbon Co). U S 1,791,558, March 3. A C black of the kind designated as "P 37" is used in making a rubber compn having an ultimate tensile strength approx the same as a similar rubber compn contg an equal wt of "common com C black" but having a stiffness materially less than such a rubber compn. An app is described suitable for making the C from hydrocarbon gas. Cf C A 24, 212.

Rubber compositions. R O COWPER. Brit 338,247, June 5, 1929 Compsns suitable for covering athletic grounds, for use as expansion joints, wall and roof coverings, machinery beds, etc., are formed of latex from which the serum constituents have not been removed, mixed with waste rubber, rubber scrap, old tire stock or the like and with breeze, ashes, clinker, brick dust, baked clay, slate dust or the like, vulcanizing materials, pigments and lustrous or portland cement, etc., may also be added, as may also fibrous fillers. The material may be placed on a backing such as cloth or paper coated with lustrous material. Various details and examples are given.

Rubber composition suitable for tire treads, shoe soles, etc. ARTHUR B. COWPER and THEODORE A BULFANT (to Barrett Co). U S 1,793,161, Feb 17. There is incorporated in a rubber mixt a coal tar substantially free from cryst material at 25° having a sp gr above 1.09 at 33° and b. above 200°. Cf. C. A 24, 1545.

Coloring rubber. I G FARREND, A G (Rudolf Kreck, inventor). Ger. 517,450, Sept 5, 1928 Rubber is mixed with an insol metal salt of Patent Blue A, which may be deposited on a substrate. The colors obtained are fast to vulcanization. Examples are given.

Coloring rubber. DUNLOP RUBBER CO., LTD., D P TWISS, R. A. MURPHY and J. G. JAMES. Brit. 338,303, Oct 11, 1929 Rubber articles obtained from aq dispersions by dipping, spreading, painting, extruding, spraying, electrophoresis impregnating or molding are colored, locally or generally, after formation, by reaction between successively added substances which together form a desired coloring substance and one of which if desired may be added to the initial dispersion, e g, β naphthol may be initially added and subsequently combined with a diazo soln, or the formed article may be dipped in indigo white soln and subsequently oxidized, or an article formed by use of acid may be dipped in an alk sulfide soln of a dye.

Preserving rubber. MARION C REED (to B F. Goodrich Co). U. S. 1,793,635, Feb 24. *Asym*-diphenyl or other diaryl hydrazine (suitably in the proportion of about 0.1-5.0%) is added to rubber compsns as a preservative.

Rubber dispersions, etc. DOUGLAS F TWISS and EDWARD A. MURPHY (to Dunlop Rubber Co., Ltd.) U S 1,793,265, Feb 17. For the production of substantially reversible compsns of pasty consistency from aq dispersions contg rubber or similar materials and a protective colloid, the dispersions are evapd in the presence of one or more substances such as glycerol, glycol or a glycol alkyl ether, having as their principal function to serve in place of water as a medium in which the protective colloids, both naturally occurring and artificially incorporated, will continue to function when the water in the dispersions is substantially removed.

Making thin rubber bathing caps. MELVIN S LOWER (to Sun Rubber Co). U. S.

1701192 Feb 24 App and various details (mainly of mech character) are described. Rubber footwear. JAMES B CROCKETT (to Cambridge Rubber Co.). U. S. 1795,075, March 3. A last is dipped into an aq rubber-contg material contg a coloring agent to form a deposited coating on the last, and the coated last is then dipped in an aq rubber compn of different coloring and the sole portion of the article is selectively dipped in an aq rubber compn contg fillers for increasing the resistance of the sole to wear.

Molded rubber footwear. H C L DUNN. Brit. 337,831, Aug 9, 1929. Articles such as galoshes or "seaside shoes" are formed by molding over a prevulcanized lining drawn onto the last and there covered with unvulcanized rubber to be molded.

Rubber battery boxes, etc. J FRAGLSON & SONS, LTD., and J. E. FRAGUSOV, Brit 338,114, July 10, 1929. Boxes are formed in such a way that the sides are harder than the base (as by subjecting the sides to a higher heat than the base in the vulcanization), they may be made of raw rubber or of waste rubber from mill tires, mixed with S, ZnO, an accelerator such as diphenylguanidine and a lubricant such as paraffin. Various details of manuf are given.

Rubber cement. THORNDON WHITFIELD. U S 1,793,583, Feb 24. Rubber latex is used with a solvent such as C_6H_6 and a small quantity of an emulsifying agent such as Na oleate, which facilitates diss of the cement with water.

Synthetic rubber. I G FARNHAM A-G. Brit 337,460, Aug 2, 1929. Alkali metal such as Na used in effecting polymerization of diolefins is enclosed in a small perforated container (which may be formed of glass or metal) placed within the reaction vessel. When the small perforated container becomes filled with polymerizate the latter with some assocd alkali metal is forced out into the main reaction vessel to promote further polymerization of its contents. The process may be carried out in 2 stages and a diluent such as diethylene dioxide may be used.

Synthetic rubber. I G FARNHAM A-G. Brit 338,152, May 14, 1929. Polymerization of diolefins such as butadiene is effected with at least 30% of an inorg O-contg acid or an org deriv of such an acid (without addn of any other substance influencing the reaction). Among the substances which may be used are: H_2SO_4 , H_3PO_4 , benzenesulfonic acid, toluenesulfonic acid, *p*-toluene sulfonylchloride and benzene sulfonylchloride. Products of varying character are obtained by varying the conditions of the reaction, and they may be suitable for adhesives, as addns to paints or lacquers, manuf of threads, films, molded articles, etc. Several examples with details of procedure are given.

Rubber-like product. JEAN BARR. Swiss 142,450, Jan 8, 1929. A rubber-like mass is obtained by treating an isoprene halide with alkali or alk. earth polysulfides. In the example, isoprene dibromide is treated with K_2S_8 . Cf. C A 24, 1545.

Factice. JEAN BARR. Swiss 142,354, to 142,357, June 10, 1928. Addns to 137,477 (C A 24, 4423). Factice sol in hydrocarbons is prepd. by heating the rubber-like mass from $\text{C}_6\text{H}_5\text{Cl}$ and a polysulfide soln with an oil such as rapeseed oil. The $\text{C}_6\text{H}_5\text{Cl}$ may be replaced by $\text{C}_6\text{H}_5\text{Br}$, $\text{C}_6\text{H}_5\text{I}$, or $\text{C}_6\text{H}_5\text{O}$.

Rubber vulcanization accelerators. ALBERT M CLIFFORD (to Goodyear Tire & Rubber Co.) U S 1,792,819, Feb 17. As an accelerator, there is used the reaction product of cyclohexylamine and an aldehyde such as aldol, butyraldehyde, crotonaldehyde, heptaldehyde or formaldehyde.

Rubber vulcanization accelerators. LORIN B. SERRALL and DEXTER N SHAW (to Goodyear Tire & Rubber Co.) U S 1,792,779, Feb 17. *o*-Tolylamidoxime or other compds of the general formula $\text{RC}(\text{NOH})\text{NH}_2$, in which R represents a benzene group, are used as accelerators.

Rubber vulcanization accelerators. JAM TEPHMA (to Goodyear Tire & Rubber Co.) U S 1,792,780, Feb 17. Vulcanization with S is effected in the presence of the reaction product of mercaptobenzoxazole and a basic amine such as diphenylguanidine. Cf. C A 24, 3135.

Apparatus for vulcanizing lengths or sheets of rubber or rubberized material. HARRY WILLSHAW and WALTER G GORHAM (to Dunlop Rubber Co., Ltd.). U. S. 1,793,269, Feb 17. Structural features

The Stroboglow. W. PINDER, BAILES AND D. D. KNOWLES. *Elec. J.* 28, 250-3 (1931).—This instrument is a rugged portable stroboscope. The true behavior of the various parts of a machine having periodic motion can usually be learned only when the machine is in normal running operation. The Stroboglow is designed to permit the study of moving parts in out-of-the-way places. It is portable, self-contained, and is packed in two large size suitcases. There are no moving parts and expert service is unnecessary. It has a broad speed range. The requisites of the ideal stroboscope are summarized as (1) ruggedness, (2) portability, (3) simple and effective control of large currents, (4) exceedingly short duration of light pulsations, (5) very intense light pulsations, (6) minimum of moving parts, (7) simplicity of operation and (8) accuracy of timing light flashes. The Stroboglow, it is claimed, fulfils these requirements. It consists essentially of the oscillator or interrupter and the light unit. The heat of the light unit is the power grid glow tube. The instrument is illustrated and the oscillator circuit shown in a diagram. Various applications are indicated. W. H. BOYNTON

Color pyrometry. GRAHARD NARSEN. *Mitt. Kaiser-Wilhelm Inst., Essenforsch. Düsseldorf* 12, 299-316(1930).—An instrument has been constructed for selecting 2 widely sepd spectral colors by means of a double prism. Light from an unknown source is then compared with light from a Pt strip by adjusting intensities on a half-shadow field. The combination of red (0.65 μ) and green (0.53 μ) was found to be the most sensitive combination for several different observers, the max error amounting to 4°. The instrument could also be used as an intensity pyrometer, and the const. in Wien's formula was measured and found to be 1.4235 cm/degree in red and 1.4179 cm/degree in green. On comparison with Pt, W and oxidized Fe, it was found that the instrument registered 60° low for Pt, 20° low for W and very nearly correct on Fe. Elimination of the comparison lamp was found to be possible, but the max error was somewhat larger. H. C. D.

Determination of color with F. Herzfeld-Hoffmann's polarization-photometer. I. B. MYTEL AND A. K. KARTASHOV. *Nauk. Zapiski Tomskogo Prom.* 10, 257-57(1930).—The app. and its use are described. V. E. BARKOW

Material requirements and material testing for heavy-duty chemical apparatus. ERNST FRANKE. *Chem. Fabrik* 1931, 133-5.—A brief review of the demands placed on construction materials in regard to aging, embrittlement, durability at high temps., etc., with remarks on methods of taking test pieces and making tests for phys. properties. J. H. MOORE

The classification of wet materials. C. H. S. TUPHOLME. *Mech. World* 88, 228-331(1930).—An outline of some of the more prominent devices for screening wet materials. E. I. S.

Chemical apparatus and their relation to the provisions of the Berufsgenossenschaft of the chemical industry and to the steam-vessel regulations. SCHNEPPLE. *Chem. Fabrik* 1931, 109-11.—A discussion is given of the lack of uniformity of regulations for steam vessels in the German states, the desirability of having national regulations, the effect of new materials and specifications and the use of autogenous and elec. welding. J. H. MOORE

Simple calculations for reciprocating compressors. JOE. IMMERSCHEIT. *Apparatebau* 43, 62-5(1931).—A table of values is given to simplify the selection of the proper size of compressor for gases. An example is worked out to show how the table is to be used. M. C. ROGERS

Transfer of heat in recuperators. L. E. THURRES AND W. BESECKE. *Heat Treating and Forging* 16, 1573-6, 1579(1930).—See C. A. 23, 3378. J. BALOZIAN

Furnace-door arches and air-cooled walls. WM. C. REINFUSS. *Heat Treating and Forging* 16, 1571-2(1930).—The construction of a double furnace (for heating for steam hammers) is briefly described, a party wall with an air passage sepd the 2 furnaces. This method of construction conserves space and gives very satisfactory results. A method of constructing furnace-door arches with tongued and grooved radial blocks is briefly indicated. J. BALOZIAN

A precision cryostat with automatic temperature regulation. R. B. SCOTT AND F. G. BRICKWEDDE. *Bur. Standards J. Research* 6, 401-10(1931).—A detailed description is given of the construction and operation of an app. that will, by means of a photo-elec. cell thermo-regulator, automatically hold any desired temp. between 0 and -170° const. to 0.001°. The temp. is easily adjusted, and the degree of refrigeration is controlled by varying the gas pressure in the double-walled bath which is surrounded by a Dewar flask contg. liquid air. J. H. MOORE

Introducing alkali or alkaline earth metals or O into sealed glass vessels (Austrian

pat 121,037) 4. Purifying inert gases in electric-discharge tubes, etc. (Austrian pat 121,056) 4. Stoneware in the chemical industry (WEINREB) 19.

APPLETON, E V. The Thermionic Valve. London Methuen. 2s 6d, net.
WUNSCH, G. Regler für Druck und Menge. Berlin R Oldenbourg 207 pp
Bound, M 13 Reviewed in *Eng Progress* 12, 96(1931)

Filter. CHRISTIAN HOLSMEYER Ger 515,853, Oct. 4, 1925 A filter for mechanically and chemically purifying liquids, especially water, contains a pad of steel wool.

Leaf filter for filtering liquids. E A ALLIOTT and MANLOVE, ALLIOTT & Co., LTD Brit 338,567, Aug 15, 1929

Leaf filter for filtering liquids. E A ALLIOTT and MANLOVE, ALLIOTT & Co., LTD Brit 338,510 Aug 15, 1929

Submerged drum filter. FINEST J SWEETLAND (to Oliver United Filters, Inc.). U S 1,796,491, March 17

Submerged continuous filter. ERNEST J SWEETLAND, JOSEPH V ZENTHOEFER and JOHN T HOYT (to Oliver United Filters, Inc.) U S 1,796,492, March 17

Air filter. DEUTSCHE LUFTFILTER BAUGES. m B 11 Ger 518,439, Sept. 8, 1928

Vacuum seal for continuous rotary vacuum filters. ROBERT O BOYKIN (to Alex Mills) U S 1,795,634, March 10

Pressure filter. N A VOSTOKOV Russ appl 71,047, May 31, 1930.

Continuously working filter press. DANIEL BALLEY Fr 635,177, Aug 20, 1929

Scraper for cleaning filter presses. V I BOSISOV Russ. appl 69,890, May 12, 1930

Separating dust, etc., from air, etc. PLATON MUNTSS REFRIGERATING SYSTEM AKTIEBOLAG Ger 518,087, June 10, 1927 See Brit 292,479 (C A 23, 1452)

Device for separating dust from air by baffling and centrifugal action. HECTOR RABEZZANA (to A C Spark Plug Co.) U S 1,795,888, March 10

Baffle-screen device for removing dust from air, etc. GLENN P HEATH U S 1,796,948, March 17

Centrifugal separator. AKTIEBOLAGET SEPARATOR. Brit 339,566, March 4, 1929, in an app for sepg liquids from solids, a layer or stream of Hg serves as a carrier of the solids Various details of construction and operation are described

Gas separator suitable for separating liquid particles from steam, etc. GRANT D READSHAW and RALPH N ROBERTSON (to Blaw-Knox Co). U. S. 1,796,435, March 17

Separator for oil and water, etc. DAVID SAMIRAN and PHILLIPS MELVILLE. Ger. 518,149, Jan 1, 1929 See Brit 322,654 (C A 24, 2043)

Calorimeter. HERMANN SANDVOSS Ger. 515,712, Nov. 20, 1928 Addn. to 491,954 (C A 24, 3403)

Pyrometer (suitable for indicating the failure of a railway or other signal lamp). C W COOKS Brit 339,200, June 11, 1929 When a lamp flame is burning properly, the expansion of metal strips in the device serves to complete an indicator circuit which is broken when the flame is extinguished or becomes too low Various details of construction are described

Apparatus and method for reducing the temperature of air by dehydration and saturation. WALTER L FLEISHER (to Cooling & Air Conditioning Corp) U. S. reissue 17,998, March 10. Reissue of original pat. No 1,749,763 (C. A 24, 2012)

Bimetallic strip thermometer suitable for ovens. THERMOSTATICS, LTD., O MILLER and F MILLER Brit 339,297, Sept 20, 1929

Automatic level regulator for liquid containers, etc. WM MONCHEV. Ger. 515,677, Feb 13, 1930

Drying apparatus for pasty or gelatinous materials, especially pigments. FRITZ HAAS Ger 518,379, Aug 3, 1929

Cascading apparatus suitable for drying crystals, cereals, etc. REGINALD V. FARNHAM U S 1,796,324, March 17. Structural details of an app for drying with gaseous media

Automatic control system for apparatus for drying materials by recirculated air. GEORGE B BAILEY U S 1,795,418, March 10 Structural features and details of a thermostatic control system are described

Rotary drum for cooling or drying chemicals or raw materials. BERNHARD SÄGBARTH Ger 515,834, July 23, 1927.

Tube-cleaning device for rotary-tube drying apparatus. MASCHINENFABRIK BLICKAU R WOLF A-G. Ger 518,380, Dec. 6, 1929.

- Means for inspecting the interior of evaporators. IJANS WOLLENBERG Ger 518,314, Mar 10, 1929
- Röntgen-ray apparatus. SIEMENS REINIGER VEIPA GFS FÜR MEDIZINISCHE TECHNIK. Brit 339,212, March 1, 1929 Structural features
- Röntgen-ray apparatus. C H F. MÜLLER A-G Brit 339,215, March 9, 1929 Structural features
- X-ray tube. SENDLINGER OPTISCHER GLASWERK G M B H (Fritz Eckert, inventor). Ger. 518,172, Aug 14, 1923 Hot-cathode x-ray tubes are made of glass contg up to 8% of rare earths, with or without TiO_2 and (or) V_2O_5 . A suitable compn. is SiO_2 70, Al_2O_3 2, MgO 3, Na_2O 20 and CaO or other rare earth 5%
- Electric-discharge lamp. PATENT-TRAUHAND-GES. FÜR ELEKTRISCHE GLÜHLAMPEN (to General Electric Co., Ltd.) Brit. 339,844, May 2, 1930 Some modifications of construction are described, relating to devices such as those of Brit. 315,391 (C. A. 24, 1550)
- Electron-emission device. JOHN W. MARDEN (to Westinghouse Lamp Co.) U. S. 1,795,730, March 10 Various structural details of sealed devices contg alkali metal Cf. C. A. 25, 622.
- Electron-emitting cathodes of discharge devices. GRAMOPHONE CO., LTD., and G. B. BAKER. Brit 339,451, Nov 14, 1929. A refractory core or oxide-coated filament is coated with a substance which on heating in the absence of appreciable O leaves a residue contg the oxide of an electron-emitting metal, such as Ca tartrate or Ba tartrate, and free C. Various details of manuf. are described.
- Electron-emitting cathodes for discharge tubes. GRAMOPHONE CO., LTD., W. F. TUDHAM and P. ELLINGTON Brit. 339,057, Oct 16, 1929 Cathodes are formed with a core or layer of Al provided with an oxidized surface having on it an electron-emitting coating which may comprise Cs, Ba or other metal deposited electrolytically or otherwise on the Al (which may have been preliminarily oxidized by making it anode in a dil. chromic acid soln.) If desired the oxidized Al may be carried on a core of other material such as W or kaolin.
- Rotary-hearth furnace. ERNEST G. DE CORIOLIS, JESSE R. MOSER and ARNOLD L. LARSEN (to Surface Combustion Corp.) U. S. 1,796,144, March 10
- Movable-hearth furnace with automatic dumping mechanism suitable for heat treatments. JOHN F. BAKER (to Westinghouse Elec. & Mfg Co.) U. S. 1,795,921, March 10 Structural features
- A two-chamber furnace with movable roof. YA. A. FRANTEVZ Russ appl 50,402, Oct. 15, 1929
- Continuous-chain furnace. SOC. ANON. DES APPAREILS DE MANUTENTION ET POURÉS STEIN Fr 695,230, Aug 30, 1929
- Refractory furnace or retort setting SOUTH METROPOLITAN GAS CO. and W. T. SEACH. Brit. 339,268, Sept. 3, 1929
- Inclined-grate furnace. ERNST VÖLCKER. Ger 518,429, July 15, 1928 Means for advancing the fuel is described.
- Step grate for low-grade or pulverulent fuel. FRANZ HOLZINGER. Austrian 121,290, Sept. 15, 1930
- Traveling grate. HERMANN KRÖNAUER Ger 518,430, Oct. 27, 1929
- Oil burner (open-dish type) OTTO SCHIEDEK and BERTHOLD OSTER Ger 518,425, Sept. 25, 1929 and 518,426, April 24, 1929
- Regenerative gas furnace. FRIEDRICH SIEMENS A-G Fr 37,175, July 2, 1929 Addn. to 591,158
- Gas burner JAY M. AXTELL U. S. 1,796,915, March 17.
- Gas burner suitable for ovens. A. H. CLUTTERBUCK and S. W. HUMPHREY. Brit 338,806, Jan 30, 1930
- Diaphragm-controlled gas-valves for heating systems. PAUL F. SHIVERS (to Minneapolis-Honeywell Regulator Co.) U. S. 1,796,544, March 17. Structural features
- Fuel-distributing means for annular or chamber kilns. MICHAEL BORN Austrian 120,993, Aug 15, 1930
- Movable feeding device for annealing ovens. A-G BROWN, BOYER & CIE Ger 518,218, Jan 11, 1927
- Rotary-tube oven for treating cork with hot gases. NICOLA MASSA. Austrian 121,253, Sept 15, 1930
- Pressure-operated valve-actuating means for blast heaters, etc. ZIMMERMANN & JANSEN G. M. B. H. Ger 518,413, Dec 17, 1929

ESSENCES ET PÉTROLES. *Fr* 37,243 Mar 16, 1929. *Addn. to 673,683* (*C. A.* 24, 2314-5). In the treatment of gases and vapors with purifying or catalytic agents a vessel of annular cross section is used having the catalytic or purifying material between 2 false bottoms, the spaces between the false bottoms and the bottoms serving resp. as admission chamber for the gases to be treated and exit chamber for the treated gases. *Cl. C. 4* 25, 112A.

Apparatus for carrying out physical and chemical reactions. I. G. FARRINGTON & G. (Hans Rahr inventor). *Ger* 514,531, Sept. 10, 1929. *Addn. to 507,033* (*C. A.* 24, 3341). The charts mentioned in *Ger* 514,003 for circulating the reacting substances pass through tubes of various lengths.

Apparatus for testing gases or liquids by thermal conductivity measurements. CHARLES ENFLEPARD, INC. *Ger* 518,377, Oct. 18, 1929. See *Brit.* 288,078 (*C. A.* 23, 351).

Apparatus for converting concentrated solutions into granulated solids. I. G. FARRINGTON & G. *Fr* 37,264, Aug. 19, 1929. *Addn. to 673,135* (*C. A.* 24, 2341). The app. consists of a rotating plate beneath a circular distributor provided with notches so as to form a circular opening between the two.

Apparatus for the distillation of solids and liquids. HENRY DUTTY. *Fr* 37,244, July 23, 1929. *Addn. to 532,349*.

Filing for reaction and washing towers, etc. FRIST BEIL. U. S. 1,794,501, March 17. A filing device is formed from a circular blank of any suitable material, folded in a crosswise reversed manner to resemble a riding saddle, the height and breadth being substantially equal.

Apparatus for liquefying chlorine by compression. KLEBS & Co. G. M. & H. (Wilhelm Lehmann, inven'tor). *Ger* 513,929, Oct. 11, 1929.

High-speed mixer for making emulsions, effecting chemical reactions, etc. GASTON S. P. DE BATHENE. *Ger* 514,200, Mar. 14, 1929. See *Brit.* 292,332 (*C. A.* 23, 1314).

Pump and high-speed emulsifier system for making emulsions such as those of bituminous materials. G. C. HICKELL and ROAD DEVELOPMENT CO., LTD. *Brit.* 338,494 Aug. 13, 1929. Various details of construction and procedure are described.

Acid-resistant apparatus. H. FARRINGER. *Brit.* 339,002, Sept. 24, 1929. Vessels or app. resistant to HNO_3 comprise an Al alloy contg. Mg, Si, etc., where exposed to acid liquid and a more resistant alloy such as Fe-Si or Cr-Ni alloy where exposed to the fumes.

Apparatus with an immersed propeller for mixing acids, etc. MASCHINENFABRIK AUGUSTO-NIKSTERG & G. *Brit.* 338,449 Sept. 9, 1929. Various structural details are described.

Float for determination of the specific gravity of accumulator electrolytes. TITAN ACCUMULATOR, LTD. Hung 101,225, Feb. 4, 1931. "Kunstwachs" is mixed with BaSO_4 to form balls having different sp. gravities. These balls are not influenced by gas bubbles as are those made of natural wax.

Apparatus for casting films. RANGETTLER MASCHINENFABRIK AUGUST KOENIG G. M. & H. *Ger* 518,007, Dec. 23, 1928.

Apparatus for making films from masses which harden in the air. RICHARD MARINOTTI. Austrian 121,078, Oct. 15, 1931.

Sheet metal articles coated with corrosion-resisting material which can be hardened by vulcanization or other treatment. P. PICK. *Brit.* 337,320, July 23, 1928. Various details of manuf. are given, mainly of mech. character, for making funnels, drums for holding chemicals, etc.

Annealing pot. OSTERREICHISCHES STEIN-SCHUCKERT-WERKE. (Erwin Häuser inven'tor). Austrian 120,850 Aug. 15, 1931.

Acetylene generator. ATTOGENWERK SEIBER G. M. & H. *Ger* 518,335, Feb. 6, 1930. *Addn. to 474,145* (*C. A.* 23, 2833).

Apparatus for distilling water by heat from the exhaust of an internal-combustion engine. H. J. WILSON. *Brit.* 338,953, Aug. 27, 1929. Structural features.

Mercury boiler. R. P. COLLISON (to British Thomson-Houston Co., Ltd.). *Brit.* 339,237, June 14, 1929.

Aluminum-coated steel soot-blower tubes for boilers, etc. WHITTEMORE, HYLBERT, WHITTEMORE & BELKNAP and H. G. C. FAIRWEATHER. *Brit.* 338,877, Aug. 21, 1929. A tubular steel core is coated on both sides with a thin layer of Al which is adapted to alloy with the steel when the tube is placed within the boiler and subjected to heat. Various details of manuf. are described.

- Safety gas valve for cutting off gas on flame extinguishment. LOUIS KREUTZ-KAMP U S 1,790,998, March 17 Structural features
- Thermostatic electric switch. HILTON I. JONES U S 1,795,400, March 10
- Thermostatic electric switch (containing mercury). THEODORE M. BURKHOLDER (to Engineering Products Corp) U S 1,795,831, March 10 Structural features
- Thermostatic control device for electric circuits. EDWARD SCHIFER U S 1,796,726, March 17 Structural features
- Thermostatic control device for electric circuits. HARRY L. BRADLEY and GUSTAV O. WILMS (Wilms to Lynde Bradley) U S 1,797,126, March 17. Structural features
- Snap-action thermostatic device suitable for controlling electric circuits. ADOLPH A. THOMAS U S 1,795,907, March 10 Structural features
- Thermostatic device suitable for controlling the flow of steam, etc. GLENN H. MEELY (to Bristol Co) U S 1,790,124, March 10 Structural features
- Thermostatic control device suitable for control of fluid fuel supply to burners. EDWARD L. FONSECA (to Wilcolator Co) U S 1,796,196, March 10 Structural features
- Thermostatic valve suitable for controlling the flow of heating fluids. FREDERICK W. JAEGER U S 1,795,860, March 10
- Thermostatic valve H. N. ROBINSON, G. W. WORTLEY and R. & A. MAIN, LTD. Brit 339,341, Sept 4, 1929

2—GENERAL AND PHYSICAL CHEMISTRY

FREDERICK L. BROWNE

- Chemists, chemistry and people. E. T. STERN *Can Chem Met* 15, 71-3 (1931)—An address W. H. BOYNTON
- Eduard Cramer. HIRSCH *Z anorg Chem* 44, 230-1 (1931)—Obituary. E. H.
- Fritz Pregl. H. LIEB *Mikrochemie* 3, 105-16 (1931)—Obituary. W. T. H.
- Edgar Fahs Smith: Provost, chemist, friend. CHARLES F. THWING *Science* 73, 331-3 (1931) E. H.
- In memoriam—Henry Leffman—1847-1930. IVOR GRIFFITH *Am. J. Pharm.* 103, 118-9 (1931). W. G. GAESSLER
- Dr. Henry Leffman—Life story. CHARLES H. LAWALL *Am. J. Pharm* 103, 113-17 (1931)—Biography with portrait. W. G. GAESSLER
- The scientific work of P. A. Leveque. OTTO FOLIV *Chem. Bull.* 18, 99-101, 121 (1931) E. H.
- Rudolf Dietz on his seventieth birthday. F. FOERSTER *Z anorg. Chem.* 44, 229-30 (1931)—A portrait is included E. H.
- The work of Marcelin Berthelot and his influence on modern chemistry. JEAN L'EPIN LEHALLEUR *Ann soc pharm chim Sao Paulo* 2, 38-44 (1931) E. H.
- Lord Melchett and his accomplishments. JAROSLAV JELINEK *Chem. Listy* 25, 34-7 (1931) FRANK MARESH
- The achievements of Josef Schneider. E. TEYROVSKY *Chem. Listy* 25, 49-53 (1931)—A brief biography is followed by a bibliography of Schneider's contributions covering all branches of chemistry FRANK MARESH
- Some scientific-instrument makers of the eighteenth century. I. H. ROBERT S. WIMPLE *J. Sci. Instruments* 7, 241-53, 273-81 (1930); *Science* 72, 208-13 E. J. C.
- Recent advances in science: Physical chemistry. R. K. SCHOFIELD *Science Progress* 25, 308-403 (1931)—Recent work on contact catalysis is reviewed. J. S. H.
- Discoveries and investigations in the older chemistry up to the phlogiston theory of Stahl. HOFFMAN *Sitzb preuss Akad Wiss* 1931, lviii-lxvi E. H.
- The activity of the general chemistry laboratory [at the University of Gand, Belgium]. F. SWARTS. *Bull. soc. chim. Belg* 39, 444-53 (1930)—A review dealing more especially with work on org. fluorides A. L. HENNE
- What conclusions must classical physics and chemistry draw from the principle of greatest simplicity? L. ZEINER *Z. Physik* 63, 706-12 (1930)—Ten postulates are given based on the principle of greatest simplicity. Classical physics must either obey these or be wrong in its fundamental suppositions. B. C. A.
- Multiple preparation of hydrogen electrodes. V. RICHARD DAMERELL. *J. Chem. Education* 7, 1664-7 (1930)—A rubber stopper was bored to take 6 electrodes, 2 reference electrodes and a H⁺-inlet tube. This was used to stopper all beakers in which the operations required by Clark's method took place. Platinization is best at 2.5 v. It is not

necessary to bubble the H gas directly on the foil, for Pt black readily absorbs the gas in soln

ANN NICHOLSON Hlad

The equilibrium between matter and energy. I. PETRYANOV. *J Russ Phys-Chem Soc*, Phys Pt 62, 241-5(1930).—The conditions of equil between matter and energy were investigated. The following equation was derived $W_T = RT \ln K_T - ((M_2 - M_1)c^2/3)$, where W_T is total work done, M_1 and M_2 are mol wts of true gases, c is the velocity of light and $K_T = -a/a_1$, where a_1 and a are the concns of gases in the system

V. VESSELOVSKY

The molecular structure of triatomic gases. II. H_2O , H_2S and N_2O . P C. MAHANTY. *Physik Z.* 32, 108-10(1931), cf *C. A.* 24, 278.—Data on dipole moments, Raman spectrum, and infra red absorption indicate that H_2S has a triangular structure like H_2O , while N_2O has a symmetrical linear form.

T. H. CHILTON

Recent advances in the interferometric determination of the form of molecules. H MARK. *Z angew Chem* 44, 125-30(1931).—Recent significant developments in the expd methods and the theory of the x ray interferometric detn of the structure of crystals are summarized. These have been of particular value in elucidating the structure of the more complex and less sym org molts. Recent applications of x ray methods and of streams of electrons to the detn of the structure of molts in gases are discussed. The results of numerous x-ray interferometric studies of simple CH_4 derivs are correlated to show the extent to which they substantiate the regular tetrahedral symmetry of CH_4 demanded by classical theory

R. H. LOWARD

Dielectric constants of liquefied gases. J C McLENNAN, R C JACOBSEN AND J O WILHELM. *Trans Roy Soc Can* [3] 24, Sect 3, 37-40(1930).—The dielec consts of liquid N_2 , A , CH_4 and N_2O were detd in a specially constructed flask and at various temps at and above their f ps. The dielec consts referred to vacuum and at selected temps are as follows: N_2 at 70.5°K 1.478, A at 87.1°K 1.620, CH_4 at 100.7°K 1.694 and N_2O at 70.8°K 2.023

J. W. SHIPLEY

The dielectric constant of supercooled sulfur and several sulfur solutions. STEFAN ROSENTHAL. *Z Physik* 66, 652-6(1930).—For the examn of liquids that attack metal plates, a glass condenser was made, consisting of concentric glass vessels. On the inside surface of the inner vessel was a sheet of Ag , on the outside of the app a sheet of Cu . The liquid was admitted between the glass vessels through connecting tubes. The dielec const of S was 3.720 ± 0.003 at 130°. Values were obtained from 135° to 95°. At 107° a sharp change was observed. Liquid S followed the law of Clausius-Mossotti (polarization, $p = \text{const}$), $p = 0.20280 \pm 0.00006$. This law was also followed by S in C_4H_8 and CS_2 .

L. P. HALL

The dielectric constant of liquid bromine. DONIESLAW DOBOSZYŃSKI. *Z. Physik* 66, 657-68, *Bull intern acad polonoise* 34, 97-111(1930), cf *C. A.* 22, 1897.—With the condenser described in the preceding abstr, the dielec const of Br was measured from 0° to 53.8°, 3.334 ± 0.010 was found at the lower, and 2.964 at the higher, temp. Polarization was a function of temp. Thus, Br followed the Debye theory for dipoles and the dipolar moment was calcd as 0.49×10^{-18} e. s. u.

L. P. HALL

Thermodynamic properties of dichlorodifluoromethane, a new refrigerant. I. Equation of state of superheated vapor. RALPH M. BUFFINGTON AND W. K. GILKEY. *Ind Eng. Chem* 23, 254-6(1931).—The consts. of a Beattie-Bridgeman equation of state for superheated CCl_2F_2 vapor were detd from measurements of isometrics between vols of 1.4 and 4.2 l per mol. In this region the isometrics are straight. The final equation is $p = \{RT(V + B)/V^2\} - (A/V^2)$, where $A = 23.71[1 - (0.305/V)]$ and $B = 0.89[1 - (0.622/V)]$. The units are atms, degrees abs and l per g.-mol. The equation fits the observed data with an av. error of 0.5%.

A. L. H.

The most important methods for determining surface tension. E. HIEDEMANN. *Z. physik chem Unterrichts* 44, 1-12(1931).—The various methods of surface tension measurement are discussed from the standpoint of their adaptability as student expts in physics labs. Four methods are suggested: (1) the torsion balance method of Lenard (cf *C. A.* 18, 2827), (2) capillary-rise method in a specially designed simple app; (3) a differential capillary rise method using an app in which 2 capillary tubes and a tube (a) of wider bore are attached to a horizontal tube and fitted with a close fitting thermometer which serves as a piston to regulate the height of the liquid in the capillary tubes, (4) a compensation method of Ferguson and Dowson (cf *C. A.* 16, 2436). Bibliography

E. R. SCHIERZ

The density of liquids. OSCAR DU SOUZA VIEIRA. *Rev brasil. chim* 1, 92-6 (1929).—A study of the methods of operating the hydrostatic balance in the detn of the d of liquids

JOHN M. LADINO

The critical anisotropic point of ferromagnetic crystals. N. S. AKULOV. *Physik*

Z. 32, 107-8(1931).—The relative change of elec. resistance of a ferromagnetic crystal under the influence of magnetization is expressed by an equation of the same form as that previously given for the change in phys. properties (C. A. 24, 2345). The following law is deduced: If 2 curves of the relative change of elec. resistance vs magnetization, derived for 2 different angles between the magnetization vector and chief axis of the crystal, intersect, then curves for any angles must also pass through this point. An example is given from data of Webster (C. A. 21, 204, 1748). T. H. CHILTON

Translation-gliding in crystals of the sodium chloride structural type. M. J. BUERGER. *Am. Mineral* 15, 174-87, 226-33(1930). cf C. A. 24, 4680.—For crystals of NaCl-like structure, the translation direction, t , is detd by the densest rows of like-charged ions. The translation plane, T , is a function both of ion radius ratio and of the polarizabilities of the constituent ions. Translation along (001) is associated with high polarizabilities, translation along (110) with low polarizabilities; with intermediate values both (001) and (110) may function as translation-gliding planes under favorable conditions of load. If the original lattice coordinates of the geometrical crystal structure theory are retained, it is found that translation direction, unit translation, direction of closest equiv. lattice points and direction of closest like charges are identical. The correlation, with minor modifications, extends to other crystals such as magnesite and other carbonates, barite, fluorite, anhydrite. It is pointed out that the translation direction is sp. for a given structural type because the translation direction has the structural character alignment of closest like leptons for all leptons in the structure. An adequate mechanistic explanation of the gliding direction is presented on the assumption that the leptons are individually charged and thus have the character of ions. The actual translation direction is the one resulting in a min. loss of cohesion.

A. M. BRANT

The mechanism of activation of thoriated tungsten and molybdenum. A. GEHRTS. *Z. tech. Physik* 12, 60-71(1931).—A math. treatment. The mechanism derived on the assumption of Th migration by slip along the crystal surfaces to the surface of W and Mo wires agrees with the exptl. results. The grain-size-diminishing action of thorium in wires may be due to adsorption and migration of Th on the crystal surfaces. A. T.

Addition centers as coordination centers. The crystal structure of veronal. ENRIQUE HERTEL. *Z. physik. Chem., Abt. B*, 11, 279-90(1930).—The crystal structure of veronal, 5,5-diethylbarbituric acid, was detd. crystallographically and roentgenographically. Crystals, m 191°, d 1.220, formed by slow growth from aq. soln., were rhombohedral, with axial ratios $a:b:c = 0.491:1:0.688$, translation group I_4' , space group V_4^{17} , 4 mols. forming the elementary group. Models are depicted. The arrangement of the mols. is attributed to secondary valences. L. P. HALL

Growth and solution of nonpolar crystals. I. N. STRANSKI. *Z. physik. Chem., Abt. B*, 11, 312-9(1930); cf C. A. 22, 4280.—Soln. of homeopolar crystals begins at the corners and edges and growth occurs more easily at the face. Metals in electrolytes, however, grow from the corners. Since homeopolar crystals present more active centers than heteropolar, they should be more effective catalytically. L. P. HALL

Explanation of the electrocapillary Becquerel-phenomenon. A correction. H. FREUNDLICH AND K. SÖLLNER. *Z. physik. Chem., Abt. A*, 152, 313-4(1931); cf C. A. 23, 4394.—The only condition necessary for the sepn. of metal in the Becquerel phenomenon is a pos. value for the decompn. potential, with the convention of a circular current going in one direction through the salt soln., in the other through the metallically conducting membrane. W. W.

Method of obtaining a single crystal of aluminum of any desired crystallographic orientation. T. SANO. *Mem. Coll. Sci. Kyoto Imp. Univ.* 13A, 307-9(1930).—An Al crystal of any desired orientation can be obtained by bringing another crystal into contact with molten Al and lowering the temp. very slowly. B. C. A.

Crystal structure of metallic lanthanum. J. C. McLENNAN AND R. W. MCKAY. *Trans. Roy. Soc. Can.* [3], 24, Sect. 3, 33-5(1930).—The x-ray spectrum of powdered La was examd. and La found to crystallize in the hexagonal system. The coordinates of the atoms referred to hexagonal axes are $\frac{1}{2}, \frac{1}{2}, \frac{1}{2}$ and $\frac{1}{2}, \frac{1}{2}, \frac{1}{2}$. The side of the unit cell is 3.72 Å. and the axial ratio is 1.63. This arrangement is the same as that of the hexagonal close packing of spheres. J. W. SHIPLEY

Crystal structure of uranium. J. C. McLENNAN AND R. W. MCKAY. *Trans. Roy. Soc. Can.* [3], 24, Sect. 3, 1-2(1930).—An x-ray analysis of the crystal structure of metallic U was made by the powder method. The calcd. spacings are for a body-centered cubic arrangement with the side of the unit cell equal to 343×10^{-8} cm. The lines corresponding to the 222 and 004 planes were not observed. Calcd. from the data obtained the d is 19.6. J. W. SHIPLEY

New forms of crystalline ice. N. ALTSBERG AND W. TROSCHEW. *Naturwissenschaften* 19, 162-4(1931)—In the Kungur caverns (Ural mountains) natural temp and satn conditions are such that the growth of large ice crystals (up to 50 cm. in size) is promoted. A few photographs of specimens are presented. B. J. C. VAN DER HOEVEN.

Structure of silicon tetrafluoride. G. NATTA. *Gazz. chim. Ital.* 60, 911-22(1930)—A survey of the literature shows that the lattice structures of very few binary compds. of the AB_4 type have been studied by x rays. It was of interest therefore to exam. compds. of this type, particularly those with a non ionic structure. The form and symmetry of the cryst. lattice of the latter depend, to a greater degree than for ionic compds., upon the form and symmetry of the individual mol., which maintains its structural unity and behaves as a fundamental constructive element of the cryst. edifice, like the single ions of ionic compds. The present paper deals only with SiF_4 , but is part of a more extended series of expts. on the halides of C, Si and Ti. Most of them give excellent x ray photographs, all of which show the cubic or tetragonal system. SiF_4 has the simplest structure. Since most of the halides have a very low m. p., the app. already used in exam. liquids and gases was again employed (cf. *C. A.* 24, 4973). Pure SiF_4 was prepd. by freezing out at -60° the HF from the gas evolved from a mixt. of K_2SiF_6 , Al silicate and concd. H_2SO_4 . The SiF_4 was then liquefied by liquid air. The powder method was used for the x ray examn. The lattice of the Si atoms in SiF_4 is body-centered cubic. The elementary cell contains 2 mols., with $a = 5.41 \pm 0.01$ A. U. and $v = 158 \times 10^{-24}$ cc. at -170° . The d is 2.17. SiF_4 is not an ionic compd., and it possesses a "mol. lattice". It belongs to the O_h spatial group. The position of the atoms in the cell is defined by the coordinates

$$\begin{aligned} &Si\ 0\ 0\ 0, \frac{1}{2}\ \frac{1}{2}\ \frac{1}{2} \\ &F\ u\ u\ u, u\ \bar{u}\ \bar{u}, \bar{u}\ u\ \bar{u}, \bar{u}\ \bar{u}\ u \\ &\frac{1}{2} + u\ \frac{1}{2} + u\ \frac{1}{2} + u, \frac{1}{2} + u\ \frac{1}{2} - u\ \frac{1}{2} - u \\ &\frac{1}{2} - u\ \frac{1}{2} + u\ \frac{1}{2} - u, \frac{1}{2} - u\ \frac{1}{2} - u\ \frac{1}{2} + u \end{aligned}$$

The closest agreement between the calcd. and exptl. intensity was found for values of u around 0.17. The radius of the neutral F atom is approx. 0.50 A. U. C. C. DAVIS.

An x-ray study of mannitol, dulcitol and mannose. THORA C. MARWICK. *Nature* 127, 11-2(1931)—Results

	Space group	a Å.	b Å.	c Å.	Density g./cc.	Molecules per cell
Mannitol	O_h	8.65	16.90	5.56	1.497	4
Dulcitol	C_{2h}	8.61	11.60	9.05	1.466	4
Mannose	O_h	7.62	18.18	5.67	1.501	4

In all three compds. the long dimension of the mol. probably corresponds to the a -axis. C. J. HUMPHREYS.

The crystallographic character of the amides of methylbutenoic acids. J. THOREAU. *Bull. sci. acad. roy. Belg.* [5], 16, 823-8(1930). *Bull. soc. chim. Belg.* 39, 412-6—The amide of tiglic acid gave orthorhombic crystals with 8 faces. The amide of angelic acid gave long crystals but could not be measured by the goniometer. The amide of α -ethylacrylic acid gave long monoclinic crystals, and that of methylethylglycolic acid gave monoclinic crystals. WALLACE R. BRODE.

Conductivity of the crystals of sodium chloride. A. K. WALTHER AND B. M. HOCHBERG. *J. Russ. Phys. Chem. Soc., Phys. Pt.* 62, 399-408(1930)—The cond. of NaCl crystals was investigated for the temp. range 25 - 800° . Polarization and its influence on the cond. were studied. The change in the cond. with temp. above 550° is attributed to the transport of electricity by Cl ions. The cond. between 25° and 550° was expressed by the equation $\delta = Ae^{-B/T}$, where δ is the sp. conductance, T the abs. temp. and A and B are consts. for a given substance and are connected with the work done in the sepn. of ions from the crystal lattice. The deviation from the above equation at temps. higher than 550° is attributed to migration of Cl ions. V. V.

Zero volumes of crystalline organic substances. W. BILTZ. *Z. Elektrochem.* 36, 815-8(1930)—A discussion of the mol. vol. and its relationship to the additivity and constitutional properties of the at. vol. Plotting the mol. vol. of homologous series against the no. of C atoms gives a value of 137 cc. for the $-CH_2$ group, and 33 cc. for the aromatic $-C_6H_5$ group. The empirical increment for the individual groups thus permits the zero vols. of the elements to be calcd. $C_{(aromatic)} = 54 =$ graphite; $C_{(aliphatic)} = 34 =$ diamond, $H = 58 = \frac{1}{2}$ H atom, $O' = 109 = \frac{1}{2}$ O atom, $O'' = 36 = \frac{1}{2}$ O atom. The zero vols. of the oxides of N are discussed. The calcd. zero vol. for the $-CH_2$ group is 150 whereas the av. observed value is 137. The energies

of decompn and the spatial requirements are roughly proportional. The exceptionally large zero vol. of ClH_2 , calcd 23.6, observed 30.65, needs to be clarified. Staudinger's study of polyoxymethylene diacetates is considered. L. L. QUILL

Relations between thermal force, thermalysis and ionic mobility in solid salts and mixed crystals. H. RINHOEN. *Z. physik. Chem. Abt. B* 11, 321-41 (1930), cf. *C. A.* 22, 3358.—Chiefly math. L. P. HALL

Some reactions that take place in the solid state. MME. MATHIEU, MATHIEU AND PAIC. *Compt. rend.* 192, 416-8 (1931).—Reactions take place readily in the solid state but they are difficult to follow unless there is a change in color or evolution of gas. The changes in x-ray spectra before and after reaction prove that the following reactions take place when the mixts. of solid powders are heated: (1) $\text{Hg} + \text{I}_2 \rightarrow \text{HgI}_2$, (2) ZnS (wurtzite) $\rightarrow \text{ZnS}$ (blende), (3) $3\text{K}_2\text{SO}_4 + \text{Na}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 \cdot 3\text{K}_2\text{SO}_4$, (4) $\text{HgSO}_4 + 2\text{HgO} \rightarrow 3\text{HgO} \cdot \text{SO}_2$, (5) $\text{HgSO}_4 + \text{Hg}_2 \rightarrow \text{HgSO}_4 \cdot \text{Hg}_2$, (6) $(\text{PhI})_2\text{CO}_2 \rightarrow \text{PhI}_2 + \text{PhCO}_2$, (7) $[\text{HgO} \cdot 2\text{SO}_2 \cdot 2\text{H}_2\text{O}] + \text{HgO} \rightarrow 2[\text{HgO} \cdot \text{SO}_2] + 2\text{H}_2\text{O}$. Reactions 1, 4 and 6 take place with increase in mol. vol., reactions 2 and 3 with decrease, and reaction 5 with no change in mol. vol. Application of thermodynamics to reactions in the solid state is not yet practicable. F. L. BROWN

Spiral-shaped cracks in the drying of precipitates. O. SCHOKKE. *Naturwissenschaften* 18, 370 (1930).— $\text{Fe}(\text{OH})_3$ was pptd. from FeCl_3 soln with NH_3 , filtered and washed and about 60 g. of the moist ppt. was placed on a glass plate and allowed to dry in the air. After 6-10 days, there was formed in the course of a few hrs. on the bottom surface a regular spiral shaped crack. This spiral formation takes place also with $\text{Al}(\text{OH})_3$. No explanation of the observed phenomenon is given. L. T. C.

Spiral formation in chemical precipitation. R. ED. LIESBOANG. *Naturwissenschaften* 18, 645-6 (1930).—Spiral cracks during the drying out of gels recently described by Belikorr (preceding abstr.) are a modification of ordinary rhythmic ring formation in ppts., etc. Its exact cause has so far not been explained. B. J. C. VAN DER HOEVEN

The density of atmospheric air and its variations. I. MOLES, M. PAVA AND T. BATURCA. *Rev. acad. cienc. Madrid* [2], 25, No. 10, 95-170 (1930).—An historic review of the studies of the physical and chem. constituents of air followed by extensive original investigations. Methods used to det. the individual factors are described in detail, and drawings are presented of the app. used. Air samples were taken from various localities. B. S. LIVING

Change of density of ethyl ether with temperature. J. MAZUR. *Nature* 127, 270 (1931), cf. *C. A.* 25, 627.—The d of very pure Et_2O increases from 0.6901 at 35° to 0.8505 at -105.4° , below which the rate of increase of d is much slower. This checks the transition temp. found in the detn. of sp. heat and dielec. const. At the m. p. -117.2° d is 0.8054. A. T. H. FISCHER

The nature of specific properties of molecular surface fields. The structure of active carbons and the inversion of the effects of adsorption and wetting heats. II. IL'IN AND I. SHIMANOV. *Z. Physik* 66, 813-8 (1930).—Exptl. data are given on an inversion of the adsorption from aq. fatty acid solns and also an inversion of the wetting heat in the homologous alk. series by two modifications of active carbon prepd. from sugars. The Debye figures of these carbons are considered. M. McMAHON

Interface potential and reactions at surfaces. I. Reduction of permanganate by charcoal. A. WASSERMANN. *Z. physik. Chem. Abt. A*, 149, 223-39 (1930).—The actions of different charcoal preps. on KMnO_4 soln were comparatively investigated. If the velocity of reduction of the KMnO_4 is a measure of the diffusion of the MnO_4 ion to the interface, the surface charge on the charcoal should influence the velocity of decompn. The effect of p_{H} on the reaction velocity in the acid region confirms this assumption. Also KMnO_4 is reduced about equally readily by normal and by activated sugar charcoal in alk. soln where both charcoal surfaces are negatively charged, whereas in acid soln, the positively charged active charcoal is the more active. The fact that the velocity of reaction in alk. soln is independent of the concn. of the alkali is explained by supposing that the C adsorbs Ott ions only in small amts. B. C. A.

Note on the adsorption theories of Frenkel and Hückel. A. GANGULI. *Z. Physik* 66, 704-7 (1930); cf. *C. A.* 24, 2666.—The adsorption equations of F. (*C. A.* 18, 3302) and H. (*Adsorption und Kapillarkondensation*, p. 172) are identified with G.'s statistical derivation of the Langmuir adsorption isotherm. CHANNING WILSON

A study of adsorption processes. II. V. N. KRISTINSKAYA. *J. Russ. Phys.-Chem. Soc.* 61, 2111-32 (1929); cf. *C. A.* 22, 9.—Prolonged sorption by C of acetic and benzole acids is not due to the formation of solid solns or to chem. reactions. It is merely an adsorption, which is conditioned by a slow diffusion of dissolved substances. Benzole and acetic acids are adsorbed by charcoal in equiv. quantities. V. V.

The adsorptive bond. H CASSEL AND F. SALOITY. *Naturwissenschaften* 19, 110-1 (1931)—The surface tension, σ , of pure Hg was measured by the method of maximal drop pressure in the absence of air and impurities and in the presence of its own vapor and several other vapors at different p and T . The σ - p isotherms allow calcn of the adsorbed amts by the Gibbs method. This procedure is an ideal one for testing the dipole theory of adsorption (Debye, *Physica* 1, 362(1921), Lorenz and Lande, *C. A.* 17, 3818; Magnus, *C. A.* 21, 844). In accordance with London and Polanyi (*C. A.* 25, 1715; and London, *C. A.* 25, 1715) the dipole moment plays only a secondary role in adsorption. H₂O vapor, between 0° and 50° and at pressures up to 62 mm Hg, had no effect at all on the surface tension of 0.1% accuracy; this means that it is not adsorbed below satn pressure. Considerable adsorption was found for H₂O with n much smaller dipole moment, more still for hexane, cyclohexane and benzene in the order mentioned (different from the order for mol refraction) which are without dipole moment. MeOH, EtOH, and PrOH are adsorbed more strongly with increasing chain length. The S-shape of the σ - p curves proves that the adsorbed state is an assocd one. Nitromethane shows the same behavior. It would seem that dipole qualities by themselves even act as a deterrent to adsorption. B. J. C. VAN DER HORST

Occlusion of hydrogen by platinum black. A. SIEVERTS AND H. BRÜNING. *Festschr. Heraeus* 1930, 97-114.—For a particular temp and pressure the sorption of H by Pt black depends on the previous treatment of the latter. The isobars exhibit a change of direction between 0° and -20°, becoming again rectilinear between -20° and -120°.

Desorption of gases from molecularly plane glass surfaces. JAMES R. CURRY, *J. Phys. Chem.* 35, 850-73(1931)—The desorption of air, CO₂, C₂H₂Cl₂, NH₃, H₂, CO, and C₂H₄ from glass surfaces was studied, in an app with a large ratio of surface vol. and a plane surface. The adsorbed gas covered only a fraction of the surface. Desorption was complete at temps below 220°.

The heat of adsorption of certain organic vapors by charcoal at 25° and 50°. J. N. PEARCE AND G. H. REED. *J. Phys. Chem.* 35, 905-14(1931); cf *C. A.* 24, 3690.—The heats of adsorption of CH₃Cl, C₂H₅Cl, CHCl₃ and CCl₄ on charcoal were detd at 25° and 50°. There does not appear to be a temp coeff of the heats of adsorption. The mol heats of adsorption increase with the no of Cl atoms in the mol. S. L.

Chemical sorption. IV. Complex sorption processes and hydrolysis. S. LIPATOV AND N. SOKOLOVA. *Z. anorg. allgem. Chem.* 192, 383-90(1930); cf *C. A.* 24, 768, 2355.—The adsorption of NaOH by salt free cellulose films was detd. Corrections were made for swelling. Na ions are adsorbed at the positions of the OH groups of the cellulose. Max adsorption shows 3 NaOH combined with each C₆H₁₀O₅. The swelling of cellulose reaches a max at 10% of NaOH. H₂O covers the cellulose in NaOH solns with a monomol film. ARTHUR FLÖRSCHER

Adsorption in solutions in relation to the dielectric properties of the solvent. I. E. HEYMANN AND E. BOYE. *Z. physik. Chem., Abt. A*, 150, 219-56(1930); cf *C. A.* 25, 13.—The relation between dielec properties of solvents and adsorption of solutes is discussed theoretically and investigated exply. Purified almost ash free wood charcoal was the adsorbent. Adsorption of BrOH, AcOH, CH₃CO₂, butyric acid and I (violet solns only) is weak from strongly polar solvents and strong from weakly polar solvents. This antitatic relation between dipole moment and adsorption is not universal, however, and is probably influenced by the mutual affinities of solute and solvent. This suggestion is supported by the observation that I is much more weakly adsorbed from brown solns in alc, Et₂O and C₆H₆, where it is solvated, than from violet solns, where there is practically no solvation. A definite and const relation between dielec. properties and adsorption occurs only in solvents that are closely related chemically, such as a homologous series of alcs. The adsorption of BrOH or picric acid in alc. soln by charcoal decreases with increasing mol polarization of the alc., i. e., with increasing deformability of the mol. A similar relation is found between adsorption and dipole moment in the adsorption of I from the series of chloromethanes. An investigation of adsorption in solvent mixts was also made, and a connection established between the adsorption curve (relating percentage adsorption to the compn of the mixt) and the mol polarization curve. In general, there are two types of behavior. Adsorption curves are linear when both components of the solvent mixt are non polar (C₆H₆—CCl₄) or when one component is polar but when its sp polarization is not altered appreciably by the presence of the non polar solvent (mixts of C₆H₆ with PhMe, CHCl₃ or PhCl). Complicated curves showing one or more minima are obtained when the sp. polarization of the polar component is altered by admixt with the non polar component (C₆H₆ with PhNO₂ or alc., CCl₄ with alc. or Me₂CO), or when both components are polar. In such

cases the max. polarization corresponds with the min. of adsorption. Adsorption in solvent mixts is further discussed in relation to vapor pressure and soly. B. C. A.

Oxide hydrates. XXXV. The crystallized oxides and oxide hydrates of aluminum as adsorbents for organic dyestuffs. GUSTAV F. HÜTTIG AND ALFRED PETER. *Kolloid-Z.* 54, 140-7 (1931), cf. *C. A.* 24, 2075, 5854.—The adsorption of methyl violet, methyl blue, Congo red, eosin and Bordeaux red R by artificial hydrargillite, boehmite, γ - Al_2O_3 (Willstätter), corundum and natural hydrargillite, bauxite and diaspor was detd. In neutral solns. the adsorption follows Freundlich's isotherm. With the natural ores the accessory minerals play an important part in the adsorption. The oxides in order of decreasing adsorption are boehmite, γ - Al_2O_3 , diaspor, artificial hydrargillite, corundum. Selective adsorption is greatest for corundum, while boehmite shows hardly any selective effect. ARTHUR FLEISCHER

Direct measurement of the adsorption of soluble substances by the bubble method. DAVID M. GANS AND WM. D. HARRIS. *J. Phys. Chem.* 35, 722-30 (1931).—The adsorption of *p*-toluidine and isoamyl alc. in the air soln. interface was measured by the bubble method. The adsorption exceeds that deduced from the equation, $u = (1/RT)(\delta\gamma/\delta \ln a)$, in agreement with measurements of McBain and Davies (cf. *C. A.* 21, 3510) although the adsorption is, in general, smaller. As the adsorbing bubbles grow smaller adsorption appears to decrease, approaching the value for the solute obtained from the adsorption equation. S. LENHER

Separation and size distribution of microscopic particles. An air analysis for fine powders. PALL S. ROLLER. *Bur. Mines, Tech. Paper* 490, 46 pp (1931), cf. *C. A.* 25, 364.—The various methods for the particle-size analysis of a microscopic powder are described. The principle, construction and operation of an improved air separator are described. Results for the sepn. of portland cement and chrome-yellow pigment in steps of 0-5, 5-10, 10-20, 20-40, 40-60 and 60-100 μ are given in detail. A closer sepn. beginning with 0-5 μ was made for gypsum and anhydrite. For soft powders only, in the 0-5 μ range, there is an attitud. of the grains. Quant. methods for correction are described. A size distribution analysis of the particles of each of the fractions made under the microscope showed that Stokes' law is closely obeyed. In each fraction there is a small overlap of a few μ due to irregularities in the shape and surface of the grains, inhomogeneity of the density of the grains, and incompleteness of fractionation. The most important particle-size properties of a powder are those that relate to the surface per unit wt. The mean diam. is defined therefore in relation to the sp. surface. The surface mean diams. for the cement powder and chrome-yellow pigment are 9.6 and 6.1 μ , resp. The dispersion is a measure of the extent of the distribution of the particle diams. about the mean diam. of the microscopic powder. The surface dispersion, related to the surface diam., is calcd. as 10.3 and 8.1 for the cement and pigment, resp. The latter is therefore twice as homogeneous as the cement. General graphic methods are described with reference to a frequency distribution of the particle diams. The method in which the area under the curve between 2 diams. gives the surface area of the particles per g. of powder between these diams. is described in detail. ALDEN H. EMERY

Triangular diagrams for the graphic representation of colloidal systems. A. DUMANSKII. *Kolloidchem. Beihfte* 31, 418-33 (1930).—By the use of triangular diagrams the effect of such peptizing agents as polyhydroxy alcs., sugars and hydroxy acids upon the reaction between FeCl_3 and NaOH or NH_4OH is shown clearly. By this means it is possible to show the range of stability of the pos. sol., the stability of the neg. sol. and the concns. at which pptn. takes place. Peptization conditions for the pptd. $\text{Fe}(\text{OH})_3$ by peptizing agents are shown well by this means. Sensitized and stabilized regions of such hydrated sols as gelatin due to the addn. of org. substances like EtOH are also represented by these diagrams. The coagulation of hydrophilic and hydrophobic colloids under changing concns. of solvents having 3 components, i. e., water, alc. and ether are shown by the same method. The diagrams also find useful application in representing emulsification and foam formation in the realm of colloids. L. H. R.

A note on the complexity of the magnetic properties of elements in the colloidal state. S. S. BHATNAGAR. *J. Indian Chem. Soc.* 7, 957-8 (1930).—Both powd. and colloidal Bi contain oxide. BiO is weakly diamagnetic or even paramagnetic. A strip of Bi tarnished by dipping in HNO_3 is less diamagnetic than the original. That colloidal metals such as Bi are less diamagnetic than crystals of the metal (see Vaidyanathan, *C. A.* 24, 3149) is at least partly due to the presence of oxides. O. T. Q.

Observations on the dielectric behavior of dispersed systems. R. FRICKS AND L. HAVESTADT. *Z. anorg. allgem. Chem.* 196, 120-8 (1931); cf. *C. A.* 24, 3183.—Changes of dielec. const. were measured for suspensions of ThO_2 hydrate, $\text{Al}(\text{OH})_3$ hydrate and

stannic acid, on adding small quantities of NaOH and HCl. The dielec. const. was decreased in each case. This decrease can be correlated with changes in the characteristics of the suspensions. The relationship is not as clear as F. and H. previously thought. The time change of the dielec. const. was measured for a gelatin soln.; the results apparently disagree with the data for stannic acid. Colloidal Au suspensions show practically the same dielec. const. as pure water. The theory of Szegvari and Wigner (*C. A.* 18, 492) is criticized. **B. J. ROSENBAUM**

Preparation of an electrolyte-free sol of manganese dioxide. **SHRIDHAR S. JOSHI AND T. SURYA NARAYAN**. *J. Indian Chem. Soc.* 7, 883-5 (1930).— MnO_2 sol was prepd by adding concd. NH_4OH drop by drop to boiling $KMnO_4$ soln. Although the sol coagulated slowly when in contact with unglazed porcelain, animal charcoal, celluloid, glass wool, filter and parchment papers, it could be successfully purified by dialysis with a parchment membrane by either of two methods. Application of a sufficiently high neg. potential to the membrane prevented coagulation until the sol became practically electrolyte free. In one expt. a pos. potential of 200 volts caused coagulation in 10 min.; zero potential allowed coagulation in 30 min.; a neg. potential of 200 v. prevented coagulation for over 13 hrs. Sols purified by dialysis at high neg. potentials were very sensitive to added electrolytes. The impure sol could also be purified by hot dialysis. If coagulation were caused by chem. reaction with the membrane substance, hot dialysis would have been impossible because of the temp. coeff. of reaction rate. The increased stability in hot dialysis is ascribed to diminution of the adsorption of colloid by the parchment. **OSCAR T. QUIMBY**

Colloidchemical studies of dye sols. II. The solution of Congo acid sols by neutral salts. **R. TANAKA**. *Kolloid Z.* 54, 156-64 (1931), cf. *C. A.* 25, 631.—The color of the soln. and the ppt. was noted on adding varying amts. of KI , KCl , $Ca(NO_3)_2$, and $CaCl_2$ soln. to Congo red sol. A change of color to red appears at the soln. max., a point independent of the pH of the soln. $CaCl_2$ showed the greatest effect. **A. T.**

Dispersoidal investigations on selenium. I. **KYOSHI JUNA**. *Bull. Chem. Soc. Japan* 6, 23-4 (1931).—Se sols are formed when Se solns. in hydrazine hydrate are poured into H_2O , C_2H_5OH , or glycerol in the presence of air, O_2 , or CO_2 but not in a N_2 atm. **ARTHUR FLEISCHER**

The effect of stirring on the rate of coagulation of gold sol. **DULVIN JONES**. *Trans. Faraday Soc.* 27, 51-8 (1931).—The effect of stirring was detd. by calcg. β in the von Smoluchowski formula from coagulation expts. with $BaCl_2$ on a Au sol made by acetone reduction. For fast coagulation, β increases with increasing rate of stirring, for slow coagulation, β decreases with increasing rate of stirring. **ARTHUR FLEISCHER**

Coagulation. **G. WIGNER**. *J. Soc. Chem. Ind.* 50, 55-62T (1931), cf. *C. A.* 22, 4311, 4312, 23, 750, 1552.—The work at Zurich on rapid and slow perikinetic coagulation of monodisperse systems, rapid perikinetic coagulation of polydisperse systems, rapid and slow orthokinetic coagulation, rapid and slow perikinetic coagulation of needle-shaped particles and the coagulation of oppositely charged colloids is reviewed. **ARTHUR FLEISCHER**

The measurement of permeability of porous alundum disks for water and oils. **HOLBROOK G. BOTSET**. *Rev. Sci. Instruments* 2, 84-95 (1931).—Lab. expts. were conducted to gain a more accurate knowledge of the flow of liquids in underground reservoirs. The app. consisted of a vertical brass tube in which a porous disk was placed. The flow through the disk was collected and measured in a graduate. The first expts. with water indicated that the decrease of the rate of flow with time was due to hydrolysis of the glass reservoir by the water in which H_2SiO_4 was formed. With kerosene the decrease in the rate depends on the presence of unsatd. hydrocarbons. This is also true for a light lubricating oil. Expts. with a crude oil showed that dissolved O_2 decreases the rate of flow. It is concluded that when both O_2 and unsatd. hydrocarbons are present the rate of flow of an oil through a porous disk decreases, but the removal of either one permits a const. rate, because clogging of the disk is caused by the oxidation products of the unsatd. hydrocarbons. Permeability measurements applied to field conditions are questionable unless it is known that the oils tested contain no unsatd. hydrocarbons or are tested in the absence of air. **ALLEN S. SMITH**

Membrane and osmosis. **L. F. A. H. SCHREINEMAKERS**. *Rec. trav. chim.* 50, 221-9 (1931), cf. *C. A.* 24, 4686.—Theoretical. When a membrane, permeable only to H_2O , separates two liquids that differ in O II' A (osmotic water action), the direction of osmosis depends upon the II'-amt. given to the membrane by the liquid. If the pressure is the same on both sides of the membrane, H_2O diffuses toward the liquid giving the membrane, the smaller W-amt. If the membrane is homogeneous, water diffuses

toward the boundary plane with the smaller H^+ amt. Through a heterogeneous membrane, however, diffusion in the opposite direction is possible. P. L. HAWKES

Influence of neutrality of solution in electrolysis, the facts and the explanation. SUMIYAMA, J. S. *Chem Ind Japan* 31, suppl. (1944) 4 (1944). A soln tends to lose its neutrality during electrolysis unless special precautions are taken. The prevailing theory is that the electrolytic potential of the diaphragm membrane, the ξ potential, alters the mobility of the ions while passing through the capillary pores of the membrane. Perm. anion paper and cationic membranes decrease the transport no. of anions and increase that of cations. With a pair of such mem. diaphragms in a compartment cell, cations are removed more rapidly than anions, and the soln becomes acid. Solns of $NaOH$ & KCl were electrolyzed in a glass container having 2 chambers sep'd by 2 similar diaphragms. The H_2O in the electrolytic chambers was not removed during an expt. The diaphragms were perm. anion or cation paper impregnated with colloidal or with charcoarized gelatin. The ξ potential of the diaphragms against $NaOH$ & KOH , KCl and HCl were studied. With a constant applied voltage the current at first increased, reached a max., then decreased. The pH of the soln decreased promptly, reached a low and nearly constant value, then increased gradually. The current time and current densities for perm. anion and for cationic, with $NaCl$, nearly coincide, although the ξ potential of cationic was 10 times that of perm. anion. When the pH of the soln is plotted against total watts consumed for expts. at different constant voltages with a given membrane the curves coincide, showing that change in pH with a given membrane depends solely upon the amt. of current passed. Current is carried by H^+ ions from the central to the anodic chamber, by H^+ ions from the central to the anodic chamber, and by OH^- ions from the cathode to the central chamber. The activity on basisity of the soln is related by the relative amt. of H^+ ions coming from the anode, and of OH^- from the cathode. The prevailing theory is justified by attributing the change in pH to two variables, based on (1) the ξ potential of the diaphragm, and (2) the concn. of ions in the diaphragm. I. M. DAVIES

Physicochemical properties of clays. I. Base exchange or ion exchange. O. WIENER, J. S. *Chem Ind* 50, 65 (1944) (1944). Minerals and clays with exchange properties are micelles consisting of an ultrathin core, an inner layer of anions, and an outer swarm of cations. The displacement of ions by exchange depends on the attractive forces between the inner and outer layer as well as upon the hydration of the ions. The ion exchange is governed by the e.c. potential of the micelle and also the behavior on coagulation. The water content and viscosity as well as the plasticity and cohesion of clay are in accord with the hydrations of the ions in the outer swarm. I. M. DAVIES

The optimum conditions for the formation of silicic acid gels from alkali silicate solutions. H. H. C. RAY AND P. H. GUNN, J. *Phys. Chem.* 48, 800 (1944), cf. C. J. 1, 24, 2158. The optimum conditions for formation of silicic acid gel by the action of HCl on Na_2SiO_3 of 1.1 N and 1.4 rather were studied. The limiting pH ranges in which gel formation takes place were noted. The large amt. of HCl in the higher pH does not affect the conditions of gel formation. The amts. of heat evolved when 1 mol. of HCl reacts with Na_2SiO_3 solns. of different ratios were also studied. The heat of reaction decreases with increase in the ratio of Na_2O/Na_2O from 1.9 to 3.4. The reaction between HCl and Na_2SiO_3 solns. results mainly in the formation of hydric silicates. H. W. EVANS

Molecularly dispersed, dissolved silicic acid and silicic acid. H. H. GUNN AND W. BRINJANON, *Chem. Ber.* 77, 41 (1944); cf. C. J. 1, 24, 11. Tetraethylorthosilicate was saponified by $NaOH$ & HCl . Data of the hydrolysis of polymer of the resulting soln. that had stood for different lengths of time made it possible (1) to find whether pure or several species of mol. of silicic acid were present after a given time, (2) to find the mol. wt. when a single species was present and (3) to follow the condensation of the silicic acid from the rapidly diffusing silicic acid, through the more slowly diffusing molecularly dispersed forms, to the colloidal ones, which no longer pass through cellophane membrane. The soln. of the freshly saponified ester contained 0.1 silicic acid, $SiO_2 \cdot 2H_2O$, mol. wt. 184. After standing 1 day this had condensed to $SiO_2 \cdot 2H_2O$, mol. wt. 294. After 1 more day the silicic acid was present as $SiO_2 \cdot 4H_2O$, mol. wt. 410. After 4 days the mol. wt. was 602. For many subsequent days the soln. seemed to contain a mixt. of differently aggregated acids. In the 45 day old soln. a silicic acid of mol. wt. 8200 seemed to be stable, but interpretation of the results for solns. of such high mol. wt. is uncertain. $NaOH$ & HCl saponified the ester more slowly than did $NaOH$ & HCl , so that drops of unsaponified ester persisted for a considerable time. Local action at the surface of the drops tended to produce nonuniformity of the species of silicic acid mol. present in the soln. However, the $NaOH$ & HCl seemed to stabilize

the molecularly dispersed silicic acid less readily than did the 0.002 N, because the change in state of mol. aggregation was more rapid. Even so, the presence of silicic acids of mol. wt. 5200 and 32,870 were noted. Orthotitanic acid triethyl ester, dissolved in MeOH, was saponified in H₂O and in 0.001 N, 0.01 N and 0.1 N HCl. The dialysis coeff. indicated that titanic acid is capable of forming molecularly dispersed acids, but that they are only slightly stable and alter rapidly to colloiddally dispersed TiO₂ hydrate. The soln. of sapon. titanic acid ester stiffens to a jelly after 6 to 22 hrs. Gelatinous silicic acid is formed as the result of the sapon. of Et orthosilicate in a mixt. of H₂O + EtOH (10:1) after several weeks, providing opportunity is afforded for the alc. to evaporate. *C. A.* 24, 1812-1911.

R. H. LOWHARD

Changes in the viscosity of gelatin sols in the process of gelation. S. N. BANERJI AND S. GHOSH. *J. Indian Chem. Soc.* 7, 923-31 (1930).—Gelatin sols of various concns. (2-10%) were prepd. and divided into 2 parts. Part I was kept at 50° and part II was allowed to gel and then kept at 25°. Parts I and II were then brought to an intermediate temp. at which viscosity measurements were made with an Ostwald viscometer. The viscosity of I increased and that of II decreased with time. Such effects are greatest near the gelation temp. Inoculation of gelatin sols near the gelation point with gel of the same concn. causes an abrupt increase in viscosity which is greater the nearer the temp. is to the gelation point.

OSCAR T. QUINCY

Simultaneous action of electrolytes and alcohol on gelatin below the isoelectric point of Michaelis. ST. PRITZLECK. *Compt. rend. soc. biol.* 106, 134-6 (1931); of *Compt. rend. soc. biol.* 104, 1079 (1930).—In a study of the action of salts and acids on gelatin sols to which alc. had been added the following results were obtained. (1) Action of acids. HCl impedes the coagulation up to a certain limit. The action of H₂SO₄ is much less marked. HCl added to the sulfate of the gelatin diminishes the H₂SO₄ coagulation. (2) (a) Action of salts on the chloride of gelatin. NaCl, MgCl₂, CaCl₂, AlCl₃ or FeCl₃ added in varying amt. to a soln. of 10 cc. of gelatin contg. 5 millimol. of HCl (*f*₉₀ 1.8) causes no opalescence. MgSO₄, Na₂SO₄ or Fe₂(SO₄)₃ causes a strong coagulation in solns. varying from 0.5 to 6 millimol. The addn. of HCl to this gelatin does not affect the coagulation phenomena. If Na₂SO₄ is replaced by H₂SO₄, merely a simple opalescence is observed. The differences between the action of Na₂SO₄ and H₂SO₄ do not depend on the differences of *f*₉₀ but on the differences of the influence of the particles of gelatin on the potential. The SO₄ ion perhaps causes this potential to fall below the minimum necessary to stabilizing the sol so that salts with multivalent cations protect the sol against the coagulation provoked by Na₂SO₄. (2) (b) Action of salts on the sulfate of gelatin. The coagulating action produced by large doses of H₂SO₄ is diminished by HCl, MgCl₂ or AlCl₃ (a peptizing action depending on the valence of the cation).

ALBERT L. RAWLINS

Thermal disaggregation of gelatin. (Röntgenographic study of its degradation.) O. GERNGROSS, O. TRIANGI AND P. KOEPPE. *Ber.* 63B, 1603-14 (1930).—Electro-osmotically purified gelatin contg. 0.13% of ash is heated in Jena glass vessels in 25% soln. of exactly isoelec. reaction for 24, 72, 75 and 336 hrs. at 100° and 48 hrs. at 121°. Possible peptide degradation is examd. by measurement of increase of the free amino groups according to Van Slyke (*C. A.* 23, 5207) by alkalimetric detn. of free carboxyl groups according to Willstätter and Waldschmidt-Leitz (*C. A.* 16, 944), and by the CH₂O no. Phys. changes are followed by measurements of the viscosity of the solns. and observations on mutarotation and the gelatinizing power of the heated solns. Even after 75 hrs. at 100° there is no sign of a chem. valency degradation by rupture of peptide linkings, whereas the gelatinizing power is completely destroyed, the viscosity reduced to 18.5%, and the mutarotation to 28.4% of the initial value of the intact gelatin. Protracted heating at 121° or exposure for 336 hrs. at 100° causes distinct, true hydrolysis. In slightly acid soln. (*f*₉₀ 1.8) much more pronounced peptide degradation is observed after 8 hrs. at 100°. On the basis of detns. of mol. wt. and x ray diagrams, the following picture of the structure and thermolysis of gelatin is drawn. Intact gelatin gels consist of extended structures that arise from extra mol. union of the gelatin particles. In aq. soln. of concn. lower than 0.5% at 20° these individual particles, according to the previous history of the preps., have a mean mol. wt. of 50,000-90,000 if the reaction is isoelec. The particles undergo an almost completely reversible distribution into mols. of mean mol. wt. of a few thousands when the solns. are heated for a short period at 100°, but, even under these conditions, an irreversible change that reduces the mol. wt. of the intact material is also obvious. Protracted ebullition causes the formation of individual mols. (mean wt. 4500), which do not recombine to aggregates when cooled. The power of forming gels or viscous, adhesive solns. is lost. It is assumed that the mols. of the intact gelatin are not united to the structures of mol. wt. 90,000 by main

valency forces, but this is not established with certainty. Reaggregation of the thermolyzed particles is observed to a certain degree if the solns are evapd to dryness and the residue is dissolved in water. B. C. A.

Swelling of cellulose in perchloric acid. K. ANDRESS AND L. REINHARDT. *Z. physik. Chem.*, Abt. A, 151, 425-32(1930), cf. *C. A.* 22, 4795.—Adsorption of HClO_4 by cellulose from aq. soln at 0° indicates that above the limiting concn, 9.3 M, an addn product is formed, $2\text{C}_6\text{H}_{10}\text{O}_5 \cdot \text{HClO}_4$, which shows a characteristic Röntgen diagram, the elementary unit having the dimensions $a = 16.5$, $b = 10.3$, $c = 10.7 \text{ \AA}$, $\beta = 93^\circ$. Below this concn the product after being washed with water gave diagrams characteristic of the native fiber. At higher concns esterification begins, and above 10.6 M the fiber is destroyed. W. L. HILL

Structure of the crystal lattice of cellulose. W. T. ASTBURY AND THORA C. MARWICK. *Nature* 127, 12-3(1931).—It is reasoned from crystallographic data for native and hydrated cellulose and sugars, that the six-atom sugar ring in the cryst. state possesses certain approx. const. linear dimensions, and that at least one of these dimensions usually corresponds to one of the axial lengths of the unit cell. The structures of native and hydrated or mercerized cellulose may differ in that the plane of the hexagonal glucose residues in native cellulose lies parallel to the a axis, while in hydrated cellulose the plane of the rings lies more nearly parallel to the c axis. C. J. HUMPHREYS

Cellobiosan and cellulose. K. HESS AND E. GARTNER. *Naturwissenschaften* 19, 180-1(1931).—Freudenberg (*C. A.* 24, 1213) claimed that the f_p depressions of AcOH solns of cryst. biosan acetate (Hess and Friese, *C. A.* 23, 1118), which is identical with acetylcellulose II (by x-ray), are caused by heat losses due to retarded crystn. of the solns. Recent expts on the velocities of crystn. indicate that Freudenberg's retardations depend on the choice of exptl. conditions (11° to 15° bath, 0.08° to 0.15° undercooling) which favor heat losses. It sufficiently great undercooling and a bath temp. equal to the "convergence" temp. are used (Raoult, Nernst, Abegg conditions) these abnormal delays do not occur. This reestablishes the validity of the depressions previously measured. B. J. C. VAN DER HOEVEN

The change in x-ray spectrum on swelling in water. J. R. KATZ AND J. C. DERKSEN. *Rec. trav. chim.* 50, 248-51(1931).—Inulin contg. ams of water varying from the equl. moisture content in satd. air down to equl. over P_2O_5 at 100° was examd. with K_α rays from Cu. A distinct crystal pattern, made up of concentric rings, was obtained. The spectra for inulin satd. with water and air-dry inulin (15% water) are identical. As the moisture content is reduced below about 8%, however, the spectrum changes, both with respect to diam. of rings and their relative intensity. The change in spectrum is therefore connected with the first portions of water taken up by dry inulin, that is, with the first bend in the S-shaped hygroscopicity curve (*C. A.* 11, 3141). Over much the greater portion of the hygroscopicity curve the spectrum remains unaltered. Conclusion, the swelling of inulin involves not only formation of shells of water around the micelles, but also an absorption of water within the crystal lattice. F. L. BROWNE

Preliminary study of the sol and the coagulum of ichthyocoll. H. G. BUNGENBERG DE JONG AND N. F. DE VRIES. *Rec. trav. chim.* 50, 238-47(1931).—The ichthyocolla (I) was obtained as a white, amorphous product by extg. the drug ichthyocoll. with hot water, centrifuging, filtering, adding enough NaOH to reduce the "alc. no." to a min., pptg., washing with alc. and drying in air. Warm water dissolves I completely. The 0.73% soln at 37° obeys Poiseuille's law. The curve of η_{sp} vs. viscosity is similar to that for gelation within the range η_{sp} 1-4 but differs markedly above η_{sp} 4. Between η_{sp} 5 and 10 the curve for I is very flat, with a min. at η_{sp} 8-9. Solns in the range η_{sp} 6-9.3 become turbid on cooling, with a max. turbidity near η_{sp} 9. The isoelec. point also lies between η_{sp} 8 and 9. Above η_{sp} 9 the viscosity increases to a relatively low max. at η_{sp} 12. The viscosity of pos. solns at η_{sp} 4.70, 3.16 and 1.51 is strongly affected by anions of electrolytes according to their valence, but is not affected materially by cations. At η_{sp} 3.16 coagulation is brought about by 10 mull.-equivs. of $\text{K}_3\text{Fe}(\text{CN})_6$. Simple coagula of I can be prepd. with alc., resorcinol, MgSO_4 , etc. Complex coagula with I as pos. component can be made with gum arabic, yeast nucleate, egg lecithin and soy bean lecithin. Coagula with I as neg. component can be made with elupein. However, complex coagulation of I with gelatin does not result in aq. medium in the η_{sp} range 5-9, where it was expected. The failure to coagulate is attributed to the flat course of the η_{sp} -viscosity curve. The drawing together of the 2 sorts of particles under the influence of elec. forces is perhaps too weak to overcome their powerful solvation. F. L. B.

Actual prevision of the solubility of substances. FREITAS MACHADO. *Rev. bras. chim.* 1, 151-69(1929).—In studying the possibilities of prevision, consideration is

Electrolyte-water systems. Distribution of the ions of two salts after diffusion through a membrane. H. BRINTZINGER AND W. BRINTZINGER. *Z. anorg. allgem. Chem.* 196, 61-4 (1931), cf. preceding abstracts.—In a soln of a binary salt the dialytic coeffs. of the anion and cation are necessarily equal (if the concn. is expressed in mols). However, if the soln. contains 2 salts, the faster anion and faster cation tend to pass through the membrane together and, in general, each ion will have a different dialytic coeff. Thus, with tenth-normal solns and a cellophane membrane, the following coeffs. were obtained for LiCl and K_2SO_4 in solns contg. a single salt, 0.739 and 0.574, resp., for Li , K , Cl and SO_4 ions in a soln. contg. all 4 ions, 0.409, 0.728, 0.882 and 0.361, resp. With 0.1 N solns and a parchment membrane, the following coeffs. were obtained for LiCl and KI in solns. contg. a single salt, 0.165 and 0.274, resp., for Li , K , Cl and I ions in a soln. contg. all four ions, 0.142, 0.301, 0.203 and 0.219, resp. O. T. Q.

Experiment to show the removal of ions in double decomposition. FRANK D. McCLELLAND. *J. Chem. Education* 7, 1579 (1930).—The reactions are carried out in an ordinary cond. cell in which the decompn. of the electrolyte is indicated by the glowing of an elec. lamp.

ANN NICHOLSON HIRD

Conductivity data of aqueous mixtures of hydrogen peroxide and organic acids. W. H. HATCHER AND M. G. STURROCK. *Can. J. Research* 4, 35-8 (1931).—Cond. changes on adding H_2O_2 to aq. solns. of formic, acetic, propionic and glycolic acids were observed. Propionic acid gave a sudden increase in cond. on adding H_2O_2 , which became const. within 3 hrs. The other 3 acids gave a sudden sharp decrease followed by a gradual increase to a const. value within 4 or 5 hrs. The const. values attained indicated the presence of peracids whose cond. is of the same order as that of the complexes supposed to be formed during the oxidation.

J. W. SIMPLEY

The conductivity of electrolytes in nitromethane. C. P. WRIGHT, D. M. MURRAY-RUST AND HAROLD HARTLEY. *J. Chem. Soc.* 1931, 199-214.—The elec. conductivities of LiCNS , NaCNS , KCNS , NH_4CNS , LiClO_4 , NH_4ClO_4 , TiClO_4 , AgClO_4 , HClO_4 , LiI , KI and Et_4N nitrate, iodide, thiocyanate and picrate were measured at 25° in 0.0001-0.002 N nitromethane (I) soln., by Fraser and Hartley's method (*C. A.* 20, 136). Of these electrolytes, only the Et_4N salts show a linear relation between Λ_s and \sqrt{c} ; the other solutes give curves represented by the empirical equation $\Lambda_s = \Lambda_0 - \alpha c^{1/2}$. Extrapolating these curves, the following Λ_0 values are obtained:

Salt	Λ_0	Salt	Λ_0	Salt	Λ_0
KI	122	AgClO_4	116.5	$\text{N}(\text{C}_2\text{H}_5)_4\text{CNS}$	119.7
LiI	117.5	TiClO_4	124.5	$\text{N}(\text{C}_2\text{H}_5)_4\text{ClO}_4$	113.4
KCNS	130	NH_4ClO_4	128.5	$\text{N}(\text{C}_2\text{H}_5)_4\text{NO}_3$	114.3
NaClO_4	122.5	(HClO_4)	127	$\text{N}(\text{C}_2\text{H}_5)_4\text{I}$	111.2
				$\text{N}(\text{C}_2\text{H}_5)_4\text{Pic}$	93.5

The observed and calcd. (from the Debye-Hückel-Onsager equation) results for the Et_4N salts agree closely, showing little tendency to ionic assocn. in soln. Those salts that deviate greatly from the ideal are probably highly assocd. This ionic assocn. is not entirely controlled by the elec. forces between the ions. HClO_4 is a strong electrolyte in I, while H_2SO_4 , HCl , HNO_3 , etc., are very weak electrolytes therein. On the addn. of H_2O to I the elec. cond. of electrolytes that deviate widely from the ideal is increased greatly. At 25°, the soly. of H_2O in I is 2%, while the soly. of I in H_2O is 10%. The d. and viscosity of I at 25° are 1.1312 and 0.00027, resp.

J. BALOZIAN

The conductivity of electrolytes in nitrobenzene. D. M. MURRAY-RUST, H. J. HADLOW AND HAROLD HARTLEY. *J. Chem. Soc.* 1931, 215-19, cf. preceding abstr.—The elec. conductivities of 0.0001-0.002 N nitrobenzene (II) solns. of Et_4N picrate and perchlorate, AgClO_4 and HClO_4 were measured at 25°. As an ionizing agent II behaves similarly to CH_3NO_2 , but very differently from hydroxylic solvents. The deviation of each of the electrolytes from the ideal is somewhat greater in II than in CH_3NO_2 , indicating that in the former the tendency for ionic assocn. is greater. Et_4N picrate and perchlorate are strong electrolytes, and give Λ_0 values that closely approx. the ideal. These 2 salts give results that agree with those calcd. from the Debye-Hückel-Onsager equation, while those for AgClO_4 do not agree. HClO_4 is an electrolyte intermediate in strength between the Et_4N and the Ag salts, being appreciably assocd. in II solns. It is a much stronger electrolyte than any other acid, approx. cond. measurements for HCl , $\text{C}_6\text{H}_5\text{SO}_3\text{H}$ and trinitrobenzoic acid showing that these are only very slightly dissoed. Extrapolation to infinite diln. of the graph of Λ_s against \sqrt{c} give the following values for Λ_0 : Et_4N picrate = 32.7; Et_4N perchlorate = 37.6; AgClO_4 = 38.4; HClO_4 = 43.

J. BALOZIAN

Influence of the strength of aromatic amines on their reactions. EDUARD ISERTEL AND KURT SCHNEIDER. *Z. physik. Chem.*, Abt. A, 151, 413-9 (1930).—Cond. curves were obtained for aniline HCl (I), α (II) and β naphthylamine (III), 4 bromo (or chloro)-1 naphthylamine (IV) and 1-bromo (or chloro) 2 naphthylamine (V) in acetone water soln., and for the same compds. with the addn. of free amine until further addns. were without effect. Accordingly, if the capacity of the amine to bind H ions be taken as a measure of its strength, II and III are strong, IV is medium strong and V is weak. The addn. reaction between amines and phenol derivatives in pure acetone as measured by a similar cond. method in conjunction with mol. wt. detns. indicate the same order of reactivity. In the addn. of phenol derivs. to amino groups in the crystal lattice the bond, $—O—N—$, is favored, while the weak ones favor the bond, $—N—C—$. Medium strong amines with medium strong phenol derivs. yield complex isomers (C. A. 21, 1802). The reactivity of amines with ethers (C. A. 25, 855) is parallel with their capacity to add H ions.

W. L. HILL

The relative strengths of bases in non-aqueous solutions. III. ERIC I. JONSSON AND JAMES R. PARTINGTON. *J. Chem. Soc.* 1931, 80-92.—The influence of dimethyl-, diphenyl-, phenylmethyl- and γ pyrone on the rate of esterification of $PbCH_3CO_2H$ in alc. HCl soln. is studied by calcg. the ratios of the equil. consts. for the formation of complex ions between the H ion and by each pyrone and by H_2O , resp., applying Lapworth's theory (C. A. 3, 738). The reaction is retarded in the order given by the following compds., H_2O , γ pyrone, diphenyl-, phenylmethyl- and dimethyl pyrone, the last named having the greatest effect. Alc. free from aldehyde, ketone and Ni^{2+} is prepd. by refluxing 96% spirit for 6 hrs. with half its vol. of well burnt lime and 2 sticks of NaOH per l., distg. into a flask contg. a lime and soda mixt., refluxing again, distg. into a dry flask the exit of which is guarded with a $CaCl_2$ tube, dropping clean, fresh Ca turnings into it, letting stand for a day, refluxing for 6 hrs., distg. off through a 12 bulb fractionating column and collecting the middle fraction. The esterification is slightly accelerated by small amts. of benzil and furan, and not affected by small amts. of dipropyl and diisopropyl ethers and glycerol.

J. BALOZIAN

The dissociation of strong electrolytes in nonaqueous solution. RUDOLF SCHROEDER. *Z. Elektrochem.* 36, 861-74 (1930).—The data on nonaq. solns. are reviewed. The Nernst-Thompson rule is valid for all solvents investigated. A more quant. relation for solvents whose dielec. const. is above 6.7 is given by the equation $1 - \alpha = 0.57 / (DA \cdot \sqrt{\epsilon} \cdot \sqrt{c})$. The behavior of the cond. found by Martin (cf. C. A. 23, 1802), which leads to the assumption of incomplete dissocn., appears to be generally true for other solvents. This action is shown by the fact that the curves for \sqrt{c} are higher than would be expected from the assumption of complete dissocn. On the other hand, the influence of dielec. const. and of ionic charge follow closely the Debye-Hückel theory. Likewise the values of the coeffs. are proportional to the square root of the concn.

JOHN R. HILL

The problem of electrolytic dissociation. KARL FREDENHAGEN. *Z. physik. Chem.*, Abt. A, 152, 321-79 (1931).—The true state of affairs does not agree with the assumptions that (a) the dissocg. compds. are composed of charged particles held together by Coulomb forces, (b) the solvents exert a dielec. action on the solutes, according to their dielec. const., (c) ions in soln. are solvated gaseous ions. Dissocn. and solv. do not run parallel to the dielec. const. of the solvent. The electrostatic theory of electrolytic dissocn. is untenable. A new theory is proposed. Electrolytic dissocn. is caused by the selective chem. affinity of the solvent's secondary valence forces for the polar components of the solute. A thermodynamic relation between dissocn. in the gaseous and liquid phases is given, by means of which liquid dissocn. may be calcd. from the dissocn. in the gas phase and the distribution no. The new theory reconciles the contrast between strong and weak electrolytes, explains the dissocn. of pure solvents and molten salts, focuses attention on solv. forces and the mechanism of electrolytic conductance.

FRANK URBAN

Simple deposition of reactive metals on noble metals. A. S. RUSSELL. *Nature* 127, 273-4 (1931).—U, Ti, W and Mo are deposited on a Hg surface, as expected according to the Nernst theory of electrode potential, when acidified solns. of their salts are shaken with Hg amalgam. The presence of the very slight concn. of metal is detected by their catalytic effect on the evolution of H_2 from H_2SO_4 solns. by pure Zn. The deposited metals are quite insol. in Hg.

ARTHUR FLEISCHER

Contact potentials. I. Measurement of the contact potential between metals and insulators. FREDRICH POLENNIK. *Z. Physik* 66, 619-31 (1930).—The potentials between certain metals and glass and quartz were detd. by an electrometer method.

A filament of glass or quartz in a clip of the metal being investigated, was held between charged plates and the deflection measured. From standardizations with known charges contact potentials were calcd for Pt, Fe, Cu, Au, Ag, Al, Mg, Zn, Pb and Sn against quartz as $+2.22, 1.99, 1.60, 1.60, 1.42, 0.93, 0.93, 0.45, 0.16$ and -0.30 , against glass $+1.15, 1.15, 0.58, 0.58, 0.14, 0.14, -0.29, -0.60$ and -1.14 v. These potentials were a function of the m p of the metals. II. The potential difference between salts and their saturated solutions. MARIANNE LEDERER *Ibid* 632-45.—An electrometer, in which quadrants cut from salts moved over satd solns permitted measurement of the charge of the double layer for pairs of salts. The contact potential for solid salt against satd soln was -0.18 ± 0.08 v for NaCl-KCl -0.04 ± 0.07 v for NaCl-NaNO₃ and -0.21 ± 0.06 v for NaNO₃-Ba(NO₃)₂. L. P. HALL

The theory of Nernst. OSCAR DE SOUZA VIEIRA *Rev brasil chim.* 1, 193-200 (1929).—Nernst's theory is outlined with particular attention to the difference of potential created between the electrode and the saline soln. In the case of ZnSO₄, with a Zn electrode the osmotic tension of the ions emitted by the electrode is added to the tension of the pre-existing ions, in order to equalize the tension of the soln of the metal. The effect of high concn. and successively decreasing concn. in a series of solns. is also discussed. The results conform to the Nernst formula for calcing the p d. J. M. L.

The propagation of flame in electric fields. I. Distortion of the flame surface. ERNEST M. GUÉNAULT and RICHARD V. WHEELER *J Chem Soc* 1931, 195-9.—The effect of an elec. field on the shape and speed of the flame-surface during its propagation is studied. The gaseous mixts (CO and air or O₂ dried over CaCl₂ and H₂SO₄), contained in a spherical explosion vessel of glass, were ignited electrically, and a photographic study (by the "snapshot" method of Ellis and Robinson, *C. A.* 19, 2747) made of the flame spread, with and without an elec. field (3000 and 5000 v/cm.). The field was produced between two parallel plates of a condenser, placed at the center of the vessel. Photographs show that the normal symmetry of the flame spread is destroyed by an elec. field. As soon as the flame touches the condenser plates the system tends to discharge itself, the elec. field being almost entirely destroyed and consequently the flame-surface suffering distortion. The photographs indicate that the movement of the flame-surface may be due to the movement of the heavy pos. ions dragging the flame-surface with them. Transverse to the field there is no change in the rate of growth of the flame, while longitudinally to it the speed of the flame is apparently increased toward the neg. plate of the condenser by the elec. field. J. BALOZIAN

Periodic process in a chemical reaction. W. T. GIBSON *Nature* 127, 13(1931).—The peculiar manner of slow oxidation of the K layer in a photoelec. cell that had developed a leak in the glass bulb contg. the material is described. Instead of oxidizing uniformly, sharply marked rings were attacked first. The rings were roughly circular with centers at the crack. These rings remained visible throughout the entire oxidation process. C. J. HUMPHREYS

Chemical reaction between metallic magnesium and salt solutions. (Quantitative Part, II.) ICHIRO IITAKA *Tetsu-to-Hagane* (*J Iron and Steel Inst Japan*) 16, 1056-63 (1930).—Metals of the alkali and alk. earth groups undergo hydrolysis in acids of concn. less than 10^{-3} N, instead of forming salts of the acid. The present paper deals with Mg in dil. HCl, the Cl⁻ concn. of which is kept at 10^{-3} N. The concn. of H⁺ was varied from 0 to 10^{-11} . Conclusions. The reaction velocity diminishes with diminution of H⁺, but does not vanish when H⁺ becomes 0. The velocity remains const. between H⁺ = 10^{-4} and 10^{-11} , being independent of H⁺. Therefore, hydrolysis must take place here, too, though some replacement of H⁺ by metallic Mg occurs, especially when the soln. is stirred. The stirring effect is explained by auto-catalytic action, forming an unknown compd. between Mg, H₂O and Cl⁻. Cf. *C. A.* 25, 863. M. KURONA

The study of chemical reactions from potential-energy diagrams. AXEL R. OLSON. *Trans. Faraday Soc.* 27, 69-76(1931).—Potential-energy diagrams for *cis trans* isomers are used to predict the effect of temp., pressure, solvents and light frequency on the stability. ARTHUR FLEISCHER

The decomposition tension of anhydrous uranyl nitrate and of anhydrous thorium nitrate. P. MISCIATELLI *Gazz chim ital* 60, 832-5(1930).—Anhyd. UO₂(NO₃)₂ was prepd., not by the method of Markétos (*Bull. soc. chim.* 11, 244(1912); *C. A.* 6, 3376) or of de Forcrand (*C. A.* 9, 1882) but by treating UO₂(NO₃)₂·6H₂O with liquid N₂O₄ by the procedure used by Guntz and Martin with Mn(NO₃)₂ and Cu(NO₃)₂ (cf. *C. A.* 4, 427). The app. is described and illustrated. According to the literature pure anhyd. Th(NO₃)₄ has not been prepd. heretofore except in an expt. by Kolb, which could not be repeated (cf. *C. A.* 8, 306). In the present work, a product contg. only 0.6% excess N₂O₄ and no H₂O was obtained. The tensions corresponding to the equil.:

$\text{Th}(\text{NO}_3)_4 \rightleftharpoons \text{ThO}_2 + 4\text{NO}_2 + \text{O}_2$ and $2\text{UO}_2(\text{NO}_3)_2 \rightleftharpoons 2\text{UO}_2 + 4\text{NO}_2$ were measured by the dynamic compensation method of Centnerszwer and Krustinson (*C. A.* 21, 689, Krustinson *Acta Univ. Latvianus* 17, 665(1928)). Following are the temps. and the pressures (mm of Hg), resp. $\text{UO}_2(\text{NO}_3)_2$, 87°, 95, 131°, 157; 141°, 119, 150°, 340, 160-340 $\text{Th}(\text{NO}_3)_4$, 124°, 82, 134°, 119, 145°, 119, 161°, 359. The same results were obtained with partially decomposed nitrates and with those containing only 4% N_2O_5 . Therefore in this thermal decomposition, basic salts are not formed and 2 phases are present in the systems at equilibrium, anhydrous nitrate and oxide. From the decomposition tensions the heats of decomposition can be calculated by the approximate formula of Nernst. For $\text{UO}_2(\text{NO}_3)_2$, Q has a mean value of 39,442 and for $\text{Th}(\text{NO}_3)_4$, a mean value of 81,823. C. C. DAVIS

Reduction of nitric acid by nitric oxide. I. Equilibrium. HANS DODE, *Z. anorg. allgem. Chem.* 195, 195-200(1931).—The equilibrium constant for the reaction $\text{H}^+ + \text{NO}_2^- + 2\text{NO} + \text{H}_2\text{O} \rightleftharpoons 3\text{HNO}_2$ was determined by measuring the pressure of NO above the system. Activities of the HNO_2 are considered. At results: K_{111} , 29.4, K_{112} , 50.8, K_{113} , 80.1. II. Electrochemical potential. *Ibid.* 201-6.—Measurements of the electromotive force of the equilibrium system, $\text{HNO}_2\text{--HNO}_2\text{--NO}$ gives as the normal potential for the reaction $\text{NO} + \text{NO}_2^- \rightleftharpoons 2\text{NO}_2^- + \text{e}$, 0.517 = 0.003 v and for the reaction $\text{NO} + 2\text{H}_2\text{O} \rightleftharpoons \text{NO}_2^- + 4\text{H}^+ + 3\text{e}$, 0.960 v. WILLIAM E. VAUGHAN

What are chain reactions? R. H. CRIST, *J. Chem. Education* 8, 604-9(1931).—Consecutive reactions that repeat their cycle many times after being initiated are called chain reactions. Several examples are given, such as the reactions between H_2 and O_2 , the formation of HCl from H_2 and Cl_2 and the thermal and photochemical formation of HBr . These require an initial process yielding an "active" substance that proceeds through a set of individually known or assumed consecutive reactions to the regeneration of the active product, which repeats the process. The termination of a chain occurs when one of the components of the set is removed permanently or suffers an energy loss sufficient to prevent its continued reaction. In case the number of chains initiated per unit time is greater than the number of those terminated the reaction may increase to explosive proportions. Chains may be initiated by a variety of methods, photochemical, particle and electron collisions, and chemical. The photochemical equivalence law affords a means of estimating the number of chains started. R. H. CARTER

A pressure stabilization process for crystalline hydrates. MAURICE AUMERAS, *Compt. rend.* 192, 229-31(1931).—The equilibrium conditions under which $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ can be kept in hydrated state for an indefinite period of time were investigated. This hydrate can be maintained in a desiccator containing H_2SO_4 , provided the concentration of the H_2SO_4 is such that its vapor pressure is between that of a saturated solution of Na_2SO_4 and the dissociation pressure of the hydrate $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$. Below 30° variations in temperature do not affect the hydrated state of the Na salt. 24.26% H_2SO_4 is the optimum concentration. M. McMAHON

The dissociation of mercuric halides. H. BRAUNE AND S. KNOKE, *Z. physik. Chem.* 152, 409-31(1931).—The dissociation of HgI_2 , HgBr_2 , and HgCl_2 was measured over a wide temperature range. Heats of formation from the equilibrium measurements were compared with the calorimeter determinations and were found to agree well. By comparing the measurements with theoretical calculations of the equilibrium, the moment of inertia of the halides was determined. On the basis of fixed assumptions as to specific heat values were obtained which for HgI_2 , HgBr_2 , and HgCl_2 are in the proportion 1.048:0.22, which is in close agreement with data estimated from gas kinetic data. The apparatus used is similar to that used with NH_4Cl (*C. A.* 22, 3570). M. McMAHON

Statistical treatment of reaction-velocity data. I. Critical review of current methods of computation. LOWELL J. REED AND EMERY J. THERIAULT, *J. Phys. Chem.* 35, 673-89(1931).—The usual methods of computing rate constants is reviewed. Unimolecularity of rate cannot be tested by the consistency of a series of K values. The least squares procedures described in chemical literature are not applicable to unimolecular data. Systems of weighting unimolecular constants are inapplicable in presence of constant errors. S. LENIER

The solution velocity of oxygen in water. IV. SUSUMU MIYAMOTO AND AKIRA NAKATA, *Bull. Chem. Soc. Japan* 6, 9-22(1931), cf. *C. A.* 25, 635.—Oxygen gas was passed into a stirred vessel containing Na_2SO_4 solution. The rate of oxidation of the Na_2SO_3 was independent of the concentration of the Na_2SO_3 and dependent only on the surface area exposed and the temperature. (The rate of stirring was constant and only pure O_2 was used.) The theory is proposed that only those moles of O_2 dissolve whose component of velocity normal to the surface is greater than a threshold value about 3.4 times the root mean square velocity. T. H. CHILTON

Kinetics of heterogeneous formate formation. GUSTAW BIRNSTEIN AND NIKOLAUS

spectra superposed. The powder method did not allow a decision as to the similarity of structure of Bi_2Se_3 and Bi_2S_3 . C. C. DAVIS

The two-component system: iron-vanadium. FRANZ WEISS AND WERNER ILLINGHAUS. *Mitt Kaiser Wilhelm Inst Eisenforsch Düsseldorf* 12, 317-22 (1930) — The liquidus solidus curves correspond to those of other investigators, showing a min at 1405° with $31\% \text{ V}$. The α mixed crystals from 0° to $100\% \text{ V}$ are cubic centered with lattice parameters that increase from $2.96 \times 10^{-8} \text{ cm}$ for 0 atom $\% \text{ V}$ to 2.89×10^{-8} for 50 atom $\% \text{ V}$ and 3.04×10^{-8} for 100 atom $\% \text{ V}$, thus departing from Vegard's law. The γ Fe field is bounded by a closed curve beginning at 1400° extending to $1.15\% \text{ V}$ and 1150° and then back to 900° . The α mixed crystals are transformed at 1234° to a homogeneous friable crystal mass that corresponds to FeV . This combination dissolves both Fe and V. The crystals are non magnetic at room temp. V at first increases the magnetic transformation temp and then decreases it after passing a max with $17\% \text{ V}$, and approaches room temp with $55\% \text{ V}$. H. C. DAVIS

Making a three-component liquid-vapor chart. JOHN HAPPEL AND JOHN GRISWOLD. *Chem Met Eng* 38, 92-3 (1931) — An alignment chart is described from which the vapor liquid equl relation may be ded for a 3-component system. A chart for benzene, toluene, and xylene at one atm is given. L. W. T. CUMMINGS

The system water-potassium nitrate-calcium nitrate. MOHAMMED ABDEL HAMID AND RAM DAS. *J Indian Chem Soc* 7, 881-2 (1930) — The isotherm was detd at 25° . The soly of KNO_3 is 27.30 and that of $\text{Ca}(\text{NO}_3)_2$ 57.98 g per 100 g soln. O. T. O

The acid-base-catalysis according to J. N. Brønsted. I. M. KOLTHOFF. *Chem Weekblad* 28, 110-21 (1930), cf *C A* 24, 3697 — A review. E. SCHOTTE

Thermal dissociation. J. KASTINSON. *Acta Univ Latviana Aim Fabulad Serya 1*, Fasc. 1, Burtuica, (in German, 63), (in Lettish 27-62) (1921) — Discuss phenomena of ThCO_3 , LiNO_3 , NaNO_3 , KNO_3 , RbNO_3 , and CsNO_3 are discussed. The isochores of liquid compds. between 200 and 900 mm approach a straight line. Single compds. with a common ion dissolved in ThCO_3 show a mol raising of the disocn temp. Nitrates dissolved in the corresponding nitrite show a strong initial rise. Binary mixts of unlimited soly show a min disocn temp dependent upon the mol compn. ANN NICHOLSON HIRD

Refrigerants show similarities (STARR) 13. Obtaining anhydrous AlCl_3 from natural Al-containing raw materials (SPRITZIN, GAVDEVA) 6.

ANGIOLANI, ARGO. *La chimica moderna. Teorie fondamentali*. Milan, Sonzogno 340 pp. L. 8.

BEEK, P. A. A. VAN DER. *Overzicht der elementaire Scheikunde met vele Vragen en Vraagstukken*. 2nd ed. Zwolle W. E. J. Tjeenk Willink. 232 pp. Fl. 2.40, bound, Fl. 3.25.

CHAPIN, WM. H. *Second Year College Chemistry*. 2nd ed., revised. New York John Wiley & Sons, Inc. 346 pp. \$3.

FERNALD, ERNEST M. *Elements of Thermodynamics*. New York McGraw-Hill Book Co., Inc. 329 pp. \$3.50.

GINSBURG, BENJAMIN. *The Adventure of Science*. New York Simon and Schuster 487 pp. \$5.

HARROW, BENJAMIN. *The Romance of Chemistry*. London Hamilton. 160 pp. 2s. 6d., net.

HOLMYARD ERIC J. *The Makers of Chemistry*. New York Oxford Univ Press. \$3.

LAURO, G. *Proprieta chimiche degli elementi e preparazione industriale dei piu comuni*. Milan A. Vallardi. 172 pp. L. 3.

MEIJERINGH, W. *Beknopt Leerboek der Scheikunde*. 8th ed. Zwolle W. E. J. Tjeenk Willink. 171 pp. Fl. 1.90, bound, Fl. 2.10.

NERST, W., AND SCHÖNFLIES, A. *Einführung in die mathematische Behandlung der Naturwissenschaften*. 11th ed., revised. Munich R. Oldenbourg 492 pp. M. 18, hcn. M. 20.

OSTWALD, WOLFGANG. *Kolloidwissenschaft, Elektrotechnik und heterogene Katalyse*. Dresden T. Steinkopff 110 pp. M. 5.

PANNAIN, ERNESTO. *Avvicinamento alla risoluzione dei problemi di chimica e di chimica fisica*. Naples R. Pionti. 235 pp. L. 30.

POSTER, ALFRED W. *Thermodynamics*. London Methuen 2s. 6d., net.

SENGERS, J. J. M. *Scheikundige Vraagstukken*. Rotterdam Nijgh & van Dijk. 41 pp. Fl. 50.

SMITH, ALEXANDER, AND HABER, FRITZ. *Praktische Übungen zur Einführung in die Chemie*. 4th ed., edited by Volkmar Kohlschütter. Karlsruhe: Braun. 140 pp. Linen, M 5 60

STEWART, OSCAR M. *Physics: A Textbook for Colleges*. Revised ed. Boston: Ginn and Co 770 pp \$4

WELD, LEROY D., AND PALMPE, FREDERIC, JR. *A Textbook of Modern Physics*. 2nd ed., revised Philadelphia Blakiston 743 pp \$3 75

3—SUBATOMIC PHENOMENA AND RADIOCHEMISTRY

W. ALBERT NOVES, JR

Recent advances in science: physics. L. F. BATES *Science Progress* 25, 383-9 (1931) —A review is given of recent work on α particles, ionic mobilities and related subjects

JOSEPH S. HEPBURN
Analytic atomic wave functions C. ZENER *Phys Rev* 36, 51-6 (1930) —Math.

BERNARD LEWIS

The specific charge of the electron according to wave mechanics FRITZ SAUTER *Naturwissenschaften* 19, 165-6 (1931) —The suggestion of Page (*C. A.* 24, 5602, *Phys Rev* 36, 1418 (1930)) explaining on the basis of wave mechanics the difference between spectroscopic and magnetic values of e/m is erroneous (cf. also Eckart *Phys Rev* 36, 1014, 1514 (1930), Uhlenbeck and Young, *Phys Rev* 36, 1721 (1930))

B. J. C. VAN DER HOEVEN

Interaction between excited and unexcited hydrogen atoms at large distances. E. C. KEMPLE AND F. F. RIEBE *Phys Rev* 36, 153-4 (1930) —Quantum mech. calcns. on the energy of different quantum states are given. The exact first-order computations of Kemple and Zener (*C. A.* 23, 4878) are extended to internuclear distances $R = 5, 6, 7, 10$

BERNARD LEWIS

Calculation of energy values. C. ECKART *Phys Rev* 36, 149-50 (1930) —Calcd. by a modification of the Ritz method for characteristic nos. in quantum dynamical problems, the results are given for the 2P and 3D states

BERNARD LEWIS

Physical and chemical principles that underlie the interpretation of novae. A. C. GIFFORD *Scientia* 49, 169-82 (1931) —A crit. review of existing theories of the formation of novae. G. finds the only satisfactory explanation to be that they arise from the grazing collision of 2 stars

J. B. AUSTIN

Stellar structure. F. A. LINDEMANN *Nature* 127, 269 (1931) —The assumption of a temp. of mc^2/k in the inside of a star is not necessary to explain the generation of energy by the annihilation of matter. In stars the important factor is the rate at which protons and electrons disappear as radiation. If an equil. state is attained matter can be annihilated in stars with interiors at temps. of 1.1×10^{10} .

A. F.

A new measurement of the temperature in the arc. ALFRED V. ENGEL AND MAX STEENBECK *Naturwissenschaften* 19, 212-3 (1931). —By detn. of the weakening of an x-ray passing through hot arc gases the d. of the gas and hence its temp. for known pressure is calcd. The x-ray intensity decreases for density δ by $I/I_0 = e^{-\alpha \delta}$ for α the (known) absorption coeff. and l the electrode distance. For air at atm. pressure the measurements were made with 1600 to 3300 v., 1 ma., Cu anode, x-radiation filtered through 7μ Al and passed through a 1.5-mm. hole in the water-cooled Cu electrode of the arc. Through an aperture in the other arc electrode the x-ray reaches a Geiger point counter ($p = 100$ mm Hg). The d. c. arc with stabilization was rectilinear inside a 5-cm. glass tube. The gas d. was found by measuring the no. of counter impulses on evacuating the cold arc tube and recording the d. at which the no. of impulses equaled that from the hot arc. With proper corrections for edge effects, etc., was found from 45,000 counted impulses a gas temp. of the 2-amp. air arc of 5250° abs. ($\approx 300^\circ$). A measurement by the carbon pencil method in the middle of the arc gave 3100° abs. The former figure is considered to be better.

B. J. C. v. D. H.

A charge exchange effect during the passage of protons through helium. R. DÖPEL *Naturwissenschaften* 19, 179-80 (1931). —A study of the passage of H canal rays through He revealed a charge exchange effect. The canal ray is deviated by passing through condensers K_1 and K_2 and subsequently its energy is measured. For a no. of protons n_1 at K_1 , n_2 at K_2 will hold $n_2 = n_1 e^{-\alpha}$ with α a measure for the active radius for charge exchange, a the condenser distance. It is assumed that the He pressure is low enough to prevent repeated charge exchanges of the neutral atoms between K_1 and K_2 . The active radius α has been detd. as a function of the proton

velocity v increases with decreasing ν and decreases again after reaching a max at about 10^4 cm. per sec. (analogous to the Ramsauer effect) B. J. C. v. d. H.

Mobility of sodium ions in hydrogen. L. B. LOEB *Phys. Rev.* 36, 152-3(1930).—The mobilities of Na^+ ions were measured in purified H_2 . It is found that within 2×10^{-12} sec. after their liberation the ions have more than twice the mobility of the normal ions BERNARD LEWIS

Electron diffraction. F. KROGER *Physik. Z.* 31, 1025-8(1930); cf. C. A. 24, 5062.—Electron diffraction patterns are sharper than x ray patterns, because the nucleus governs the scattering of electrons, and the electrons affect the scattering of x rays. In electron diffraction patterns the 2nd-order spectra due to the (111) planes of NaF are not entirely extinguished, since the nuclear charges of Na^+ and F^- are unequal. Debye-Scherrer diagrams due to electrons diffracted by comparatively large crystals sometimes show interference points at the ends of a diam. of one or more of the rings. These can be ascribed to a mosaic like structure of the crystal. Sublimation under vacuum of HgCl_2 from a heated Pt strip to a celluloid surface gives a cubic instead of a rhombic crystal. Diffraction by gaseous N , O or CO_2 gave patterns in accord with Debye x ray patterns A. L. HEWY

Electron diffraction and molecular structure. R. WIERL *Physik. Z.* 31, 1028 (1930).—This is a detn. of the distance between at. nuclei in a mol as measured by electron diffraction. In CS_2 , the distance S-S is 3.20 A. U.; in CO_2 , O-O is 2.26 A. U.; in N_2O , N-N is 2.38 A. U.; in SO_2 , the distance S-O is doubtful, on account of the triangular shape of the mol; the value 1.37 A. U. is proposed. A. L. HEWY

Emission of electrons caused by an electric field when the surface of a metal is exposed to the light. L. ROZENKOVICH, *J. Russ. Phys.-Chem. Soc., Phys. Pt.*, 62, 229-30(1930).—It is pointed out that a study of the emission of electrons caused by an elec. field when a metal is exposed to the light may lead to a new conception of electronic energy V. VESSELOVSKY

The photoelectric effect from sodium chloride in various gases. N. GUDRIS AND L. KULIKOVA, *J. Russ. Phys.-Chem. Soc., Phys. Pt.*, 62, 247-50(1930).—The photoelec. effect from dielectrics is highly stimulated by the presence of the vapors with high dielec. const. in the vicinity of the surface investigated. It seems that the mols. of such gases having a high dipolar moment may disengage ions from the surface of the crystal. When the vapors dissolve the crystal (water vapor) this can also take place inside the lattice, which would also increase the intensity of the photoelec. current. In the other gases the photoelec. effect is similar to that in air V. VESSELOVSKY

Photoelectric effect and electron reflection at hydrogenated potassium surfaces. W. KILGE AND E. RUPP *Physik. Z.* 32, 163-72(1931).—An app. is described which permits photoelec. and electron reflection measurements to be made on the same K surface. The surfaces studied were pure K, K in H_2 atm. (I), K subjected to weak glow discharge in H_2 (II), and K subjected to strong glow discharge in H_2 (III). The colors of these layers were: K, metallic; I, metallic; II, reddish violet; and III, greenish. With the pure K surface, the voltages causing max. electron reflection corresponded to planes in the K lattice only. With I only electron reflection max. for KH were observed. With II and III max. for both K and KH were observed, those for KH predominating in the case of II. In all cases the photoelec. effect showed a pronounced max. in the region 4000-4500 A. U. This "selective" effect increased in the order K, I, III, II. The electron reflection max. led to a value 7.3 v. for the inner potential of K and 5.4 A. U. for the lattice const. of KH (assumed cubic). O. T. Q.

A thermal theory of cathodic disintegration. H. P. WARAN *Phil. Mag.* [7], 11, 337-404(1931), cf. C. A. 16, 1043, 17, 3543.—The sputtering of cathodes is due to the evapn. of small masses of metal under localized intense heat due to proton bombardment. The observed relative sputtering of various metals agrees qualitatively with those expected from the thermal conductivities of the metal and gas surrounding the cathode, the sp. heat of the metal and the gas, the gas pressure, metal d., b. p. and latent heat of the metal. ARTHUR FLEISCHER

Measurement of capacity for reflection of glowing tungsten in the ultra-violet part of the spectrum. E. SPILLER *Z. Physik* 64, 39-42(1930).—The relative intensities of emission at different wave lengths from a lamp with a W filament heated to 2873° abs. were found, a special microphotometer being used. The emission departs from black-body radiation. The energy distribution was measured over a wave length range of 400-340 m μ . The values of the emission coeff. for wave lengths from 435 to 270 m μ for a filament at 2100° abs. are tabulated. B. C. A.

Photo-ionization of salt vapors. A. TEREVICH *Phys. Rev.* 36, 147-8(1930).—The photo-ionization of the halides of Th was investigated with wave lengths longer than

in the study of the relative reflecting power of the 2 crystals in a given spectral region
R. H. LOMBARD

The construction of an experimental x-ray tube to furnish the homogeneous K radiations of uranium. WARFORD MOPPETT. *J. Cancer Research Comm Univ Sydney* 2, 147-52(1930)—Constructional details of the app are given and the method of operating it is described.
G. SCHWOCH

Partial absorption of x-rays. B B RAY. *Z Physik* 66, 261-8(1930)—The Compton effect is studied by means of $\text{CuK}\alpha$, $\text{NiK}\alpha$, $\text{NiK}\beta$, $\text{WK}\alpha$ and $\text{FeK}\beta$ radiation sent through C, N and O. A possible mechanism for the Raman effect is worked out at some length.
GREGG M. EVANS

The paramagnetism of polychromates. FRANCIS W. GRAY AND JOHN DAKERS. *Phil Mag* [7], 11, 297-314(1931), cf *C A* 24, 5613—The mol susceptibilities were detd. in the app of Gray and Dakers (cf *Phil Mag* [7], 9, 81). The values ($\times 10^6$) are $\text{K}_2\text{Cr}_2\text{O}_7$, 0.0; $\text{K}_2\text{Cr}_2\text{O}_7 + 54.0$; $\text{K}_2\text{Cr}_2\text{O}_7 + 95.0$; $\text{K}_2\text{Cr}_2\text{O}_7 + 193.0$; $\text{CrO}_3 + 41$; $\text{Na}_2\text{Cr}_2\text{O}_7 + 19.4$. The χ susceptibility of CrO_3 in solns. with 1 mol. CrO_3 to from 39.27 to 18.89 mols H_2O was +0.39. A small paramagnetism due to the interchange of electrons and consequent fluctuation of angular momentum gave the best interpretation of the exptl results.
ARTHUR FLEISCHER

A synchronous film drum for recording periodic spectra. DONALD C. STOCKBARGER AND CARL G. SELIG. *Rev Sci Instruments* 2, 211-6(1931).
E. H.

Studies in peculiar stellar spectra. I. The manganese lines in α Andromedae. W. W. MORGAN. *Astrophys J* 73, 104-17(1931)—The spectrum of α Andromedae is of Class A_p, the peculiarity being due to the presence of many lines of ionized Mn. An examn. at the Yerkes Observatory of spectrograms of stars belonging to Classes B₁ to A₉ has revealed the presence of enhanced Mn lines in 13 other stars. The lines probably belong to Mn III, indicating that the element is present in the 2nd stage of ionization. If the high abs. magnitudes of these stars are taken into consideration the presence of enhanced Mn lines seems to be a normal feature of their spectra, the lines reaching their max. in stars of Class B₁.
C. C. KRESS

Influence of an inhomogeneous electrical field upon the fine structure of hydrogen-like atoms. SISIRENDU GUPTA. *Z Physik* 66, 246-56(1930)—The frequencies and intensities of the lines for the inhomogeneous Stark effect are derived math. and the sepn. is found to be of the order of 0.001 A. U., even for a field gradient of 10^{-4} to 10^7 . This sepn. is so small as to be discernible only as a slight broadening of the lines.
GREGG M. EVANS

The radiation emitted by compressed substances under high potentials. ISAY A. BALENEN. *Phil Mag* [7], 11, 315-22(1931), cf *Reboul*, *C A* 21, 20—The expts. of Reboul were repeated for compressed potash alum, NaCl and other powders. The factors controlling the emission were the applied potential, relative humidity, size of cryst. grains and the pressure used in compressing the powders. The very absorbable radiation is due to the escape of air from the bottom of the cell and belongs to the same class as "Entladungstrahlen."
ARTHUR FLEISCHER

Experimental evidence for the existence of quadrupole radiation. R. FRERICHS AND J. S. CAMPBELL. *Phys Rev* 36, 151-2(1930)—The green auroral line at 5577 A. U. was found to be resolved into 4 components by means of the Zeeman effect. This indicates that the line is quadrupole radiation.
BERNARD LEWIS

Regularities in the second spectrum of xenon. C. J. HUMPHREYS, T. L. DE BRUIN AND W. F. MEGGERS. *Bur Standards J Research* 6, 287-93(1931)—The ground doublet ($2P^{1/2}, 3/2$) in the spectrum of ionized Xe (Xe II) has been identified, the levels are sepd. by 10,540 wave nos. This spectral term has been connected with others which combine so as to give most of the intense lines in the visible Xe II spectrum. The abs. values of the terms have been fixed by extrapolating series forming terms to their limits, and from these spectroscopic data the ionization potential of Xe^+ is found to be about 21.1 v.
W. F. MEGGERS

A study of stellar hydrogen lines and their relation to the Stark effect. C. T. ELVEY AND O. STRUVE. *Astrophys J* 72, 277-300(1930), cf *C. A.* 24, 2950—The contours of the H absorption lines in early type stars have been detd. from microphotometer measurements of stellar spectrograms. The lines exhibit extensive wings which are evidence for the existence of a Stark effect. The theoretical contour of a line subject to an elec. field in an atm. in which the pressure increases with depth is detd. and is found to be very similar to those actually observed.
C. C. KRESS

Probabilities of recombination into the 1^3S state of cesium. C. BOECKNER. *Bur Standards J Research* 6, 277-85(1931)—Measurements are made of the intensity distribution in the continuous emission band, appearing at the 1^3S series limit of Cs, a

low-voltage thermionic discharge in the vapor is used as a source. The wave-length range covered lies between the series limit 3184 Å U. and about 2750 Å U. The methods of photographic densitometry are employed, a W-strip lamp in quartz being used as a comparison source. From the variation of intensity with wave length and the velocity distribution of the discharge electrons are computed the relative probabilities of recombination of free electrons into the 1^1S state as a function of their initial velocities. Previous work has shown that the probability of recombination into the 2^3P and 3^3D states of Cs varies approx. inversely as the square of the velocity of the free electron, a relation which is also predicted by quantum mechanics for recombination into any level of a H atom. The present measurements show that the probability of recombination into the 1^1S Cs level falls off much more rapidly with the velocity, for electron energies greater than 0.15 v. more nearly as the inverse 4th power. W. F. M.

Structure of the iron spectrum. M. A. CATALÁN. *Anales soc. españ. fis. quim.* 28, 1239-385 (1930).—Previous data on the structure of Fe are summarized and new advances given. More than 2350 lines are classified in the spectrum of Fe I, in 304 levels. The most intense lines originate by addition of a valence electron to the residue in one or both spark configurations $3d^4s$ and $3d^7$. Lines produced by a double jump are frequently intense. E. M. SYMMES.

The rarest ultimes in the spectrum of rhenium. HERBERT SCHÖBER AND JOSEF BIRKE. *Naturwissenschaften* 19, 211-2 (1931).—The rarest ultimes of the Re arc spectrum as observed from varying concns. of Re salt on C, i. e., those lines that originally have the largest intensity, are at 3452.02, 3460.61 and 3464.87 Å U. Intensity and wave length are given for 23 other Re lines. B. J. C. VAN DER HOEVEN.

Resonance radiation of manganese vapor. J. FRIDRICHSON. *Z. Physik* 64, 43-7 (1930).—The fluorescence in said Mn vapor at 700-900° was examined, the light of the Mn spark being used for excitation. The triplets $1S-2P$ and $1S-2P'$ of the Mn spectrum correspond with the resonance lines. With rise of temp., and consequently less dense vapor, the triplet $1S-2P'$ becomes much more intense, whereas at a lower temp. only the triplet $1S-2P$ is found. The resonance lines found are in agreement with Catalan's analysis (C. A. 17, 1586, 22, 1901) of the Mn spectrum and indicate that $1S$ is the lowest state of the Mn atom. B. C. A.

Spectrum of the tungsten mercury arc. W. D. FORSYTHE AND M. A. EASLEY. *Phys. Rev.* 36, 150-1 (1930).—The pressure of Hg vapor in a Hg arc lamp is increased by external heating between W electrodes and the effect on the absorption bands in the region 2100-4000 Å U. is described. The absorption spectrum is compared with that of Hg in another quartz absorption cell. BERNARD LEWIS.

Note on the spark spectra of chlorine. KIYOSHI MURAKAWA. *Sci. Papers Inst. Phys. Chem. Research (Tokyo)* 15, 105-9 (1931).—About 30 lines between 2469 Å U. and 4695 Å U. are classified as members of the quartet system of Cl III. A few corrections and extensions of the analysis of Cl II, previously reported (C. A. 25, 1157), are noted. C. C. KIESS.

Spectrum of doubly ionized iodine. J. B. SETH. *Nature* 127, 165 (1931).—A short table is given of some new measurements on doubly ionized I in the visible and the ultra-violet regions. MALCOLM DOLE.

Temperature classification of the stronger lines of columbium, with preliminary notes on their hyperfine structure. ARTHUR S. KING. *Astrophys. J.* 73, 13-25 (1931).—The spectra of Cb, as excited in arcs and in spark discharges, and in the vacuum electric furnace at temps. ranging from 2500° to 2900°, were examined in the region from 3094 Å U. in the ultra-violet to 6938 Å U. in the red. Temp. classifications have been determined for 646 lines in this interval, of which approx. 200 appear with furnace excitation. The others, consisting of lines of both the neutral and ionized atoms, require the higher excitation energy of the arc and spark for their appearance. About 40% of the lines examined are complex, the hyperfine structure consisting of 2, 4, 6, and in some cases, more components. C. C. KIESS.

Calorimetric and electric measurements in the mercury arc. JÓZSEF KAWA. *Spravoždania i Prace Polskiego Towarzystwa Fiz.* 4, 27-58; *Chem. Zentr.* 1929, 11, 2979.—The Fe anode of a Hg-arc lamp of Pyrex glass was constructed in such a way as to serve as a calorimetric vessel, which was filled with H_2O . The Hg cathode and the glass wall of the lamp up to the anode were immersed in water in order to obtain const. temp. and const. Hg-vapor tension during the arcing period. The calorimetric measurements carried out with various c.d.s. and arc lengths gave the following results with regard to the mechanism of the Hg arc. Ionization occurs in the 1st part of the positive column on account of the greater field strength. The pos. and neg. space charges are formed. Between these originates a reverse elec. field which brakes the speed of

the electrons accelerated in the 1st part of the column. Because of this also the heat formation at the anode is diminished. Between the pos space charge at the front of the pos column and the neg space charge in the Faraday dark space an increased pos. field strength is developed. With further increase of the arc length, under suitable conditions there occurs a 2nd max. These results are supported by spectroscopic investigations. The arc was mixed with some He, whose lines appear only in places of higher field strength on account of the higher excitation potential. G. SCHWOCZ.

Different spark spectra of mercury. R. RICARD. *Compt rend* 192, 618-20 (1931).—A no. of new lines have been found in the spectrum of the electrodeless discharge in Hg. These lines fall into the 3 spectra, Hg II, Hg III and Hg IV. A no. of these lines in the Hg II spectrum can be interpreted by admitting the existence of doublets and quadruplets. J. R. ADAMS.

Influence of pressure and temperature of added neon on the absorption in excited mercury vapor. OSAMI MASAKI. *J. Physik* 66, 229-40 (1920).—Ne purified by passing over liquid air and through discharge tubes with Mg electrodes was circulated through an emission and an absorption tube contg. Hg excited by 10 ma. at 8000 v. The beam from the emission tube was split, one beam passed through the absorption tube mounted in a spectrophotometer, and the absorption measured. The $2^1P_{1/2} \rightarrow 2^1S$ transitions were studied. The absorption of lines 4047 and 5461 Å. U diminishes slightly with increasing pressure. 4358 shows a slight increase. A satn is shown with increase in temp., 4358 and 4047 increasing toward about the same value, while 5461 decreases to reach a satn value slightly under that of the others. GREGG M. EVANS.

The crystal structure of parahydrogen at liquid-helium temperatures. W. H. Keesom, J. DE SMEDT and H. H. MOOR. *Nature* 126, 757 (1930).—See C. A. 25, 873. A. L. HENVE.

Spectroscopic isotope determinations. R. SCHMID. *Magyar Chem. Folyóirat* 36, 185-91 (1930).—General description of present theory and methods. S. S. DE G.

The interpretation of molecular spectra. V. The excited electron terms of molecules with two equal nuclei (H, He, Li, N, O). F. HUND. *Z. Physik* 63, 719-51 (1930), cf C. A. 21, 3310. 23, 564. W. F. MEGGERS.

A new band system of hydrogen and the conditions for its production. Z. BAY, W. FINKELNBURG and W. STEINER. *Z. Physik Chem., Abt. B*, 11, 351-62 (1931).—A strong condensed discharge ($\frac{1}{2}$ mf. capacity) produces a yellow band group of the many lined spectrum and a no. of other strong lines, while the remaining many-lined spectrum vanishes. From the variation of intensity with gas concn in the discharge tube it is concluded that the yellow group belongs to the H, mol. The group was not completely analyzed, because the weak lines were missing, but it appears to be 2-band systems with a common initial or final term. This term is not one of the known H, terms. The electron jump has the values 2.1 and 2.6 v., which excludes a transition between the 2nd, 3rd and 4th levels. J. B. AUSTIN.

Some relationships between the continuous and the many-lined spectrum of hydrogen. L. YOSIIDA HUKUMOTO. *Science Repts. Tôhoku Imp. Univ.* 19, 773-82 (1931).—The behavior of the H, spectrum under different conditions of pressure was observed and the relative intensities were estd. Most of the lines which are enhanced at high pressure when the continuous spectrum is intense are classified as the transitions $3^1\Pi \rightarrow 2^1\Sigma$, $4^1\Pi \rightarrow 2^1\Sigma$, etc. These exptl facts favor the theory that the continuous spectrum arises in the transition from any one of the excited triplet levels to the unquantized ground state $1^1\Sigma$. W. F. MEGGERS.

Nitric oxide δ -bands. R. SCHMID. *Z. Physik* 64, 279-85 (1930).—In order to produce the δ -bands of NO in sufficient intensity to be photographed special means of excitation must be used. With a Tesla app. and other devices for special excitation it was found possible to produce the bands in the afterglow in the production of active N mixed with air. The rotation analysis of the band $\delta(0.3)$ is given, and the rotation const. of the δ -bands for the $^1\Sigma$ initial level is calcd to be $B_0 = 1.995 \text{ cm}^{-1}$. Also in *Matematik és Természettudományi Értesető* 47, 534-42 (1930). B. C. A.

Nitric oxide β - and γ -bands. R. SCHMID, D. FAREAS and T. KÖNIG. *Z. Physik* 64, 84-120 (1930).—The NO lines previously measured were classified and ordered. In this way some 3000 γ - and 1000 β -lines were arranged. The rotation and vibration consts. were in some cases calcd afresh. The doublet splitting of the levels $^1\Pi_2$ was obtained as a function of rotational and vibrational states. The potential-energy curve for NO was obtained, as well as Condon's parabola. Also in *Matematik és Természettudományi Értesető* 47, 485-532 (1930). B. C. A.

The structure of the spectra of sulfur. Contribution to the study of the relations between spectra of the same multiplicities. M. GILLES. *Ann. Phys.* 15, 267-408

(1931)—Grating and prism spectrographs were employed to observe the spectra of singly and doubly ionized S from 2089 Å. U. in the ultra-violet to 7634 Å. U. in the red. The Zeeman effects of some of these lines have also been observed with fields of 31,700 gauss. These new observations together with other available wave-length data have permitted an analysis of S II into doublets and quartets, and of S III into triplets and quintets. For the most part the spectral structures which have been established in these spectra conform to the quantum-theoretical requirements. A few exceptions are noted. The terms of S III are compared with the corresponding terms of Si I and P II by means of the irregular-doublet law, similarly those of S II are compared with Cl III. Certain numerical relations between the septa of the terms of the isoelectronic systems N II and O III, and P II and S III are shown to correspond to the nuclear charges of the resp. ions. C. C. KIRBY

High-frequency discharges in nitrogen in the presence of mercury. R. ZORRER, *MANN. Compt. rend.* 192, 49-51 (1931).—The spectra of N in the presence of Hg were obtained by excitation with high-frequency currents ranging in wave length from 10 m. to 570 m. The spectra were principally the bands of N_2 , N_2^+ and CN, and numerous lines of at. Hg. The neg. bands of N_2 corresponding to low vibrational quantum nos. were very feeble for high-frequency currents but became very intense for currents of frequencies corresponding to wave length 570 m. The neg. bands of high quantum nos. were apparently unaffected by changes in the exciting currents. The pos. bands always appeared strongly. Likewise the Hg lines representing the $P \rightarrow S$ transitions were unaffected, whereas those representing $P \rightarrow D$ transitions were more intense for the higher frequency currents. C. C. KIRBY

The band spectra of scandium, yttrium and lanthanum monoxides. WM. F. MCGEE AND JOHN A. WHEELER, *Bur. Standards J. Research* 6, 509-73 (1931).—The ordinary arc spectra of Sc, Y and La show, in addition to lines characteristic of neutral and ionized atoms, complex banded structures which are ascribed to a mol. compd. of the atoms with O. These bands have the same general appearance as those due to well known diatomic mols. and are, therefore, attributed to ScO , YO and LaO , rather than to the chemically stable oxides Sc_2O_3 , Y_2O_3 and La_2O_3 . The band spectra of these oxides are so complex that the fine structure of the individual bands has not been sufficiently well resolved to permit analysis of the rotational energies, consequently only a description of band heads and their classification in various systems resulting from transitions between a no. of vibration levels of initial and final electronic states can be given. The ScO spectrograms showed 159 band heads which have been divided into 5 systems, the 0, 0 transitions being represented by band heads at I, 4837.79 Å. U., 4838.09 Å. U.; II, 6017.07 Å. U.; III, 6036.17 Å. U.; IV, 6064.31 Å. U.; V, 6079.20 Å. U. The YO spectrum closely resembles that of ScO ; 120 bands have been described for YO and divided into 5 systems with 0, 0 transitions at I, 4817.23 Å. U., 4818.20 Å. U.; II, 5909.08 Å. U.; III, 5972.04 Å. U.; IV, 6036.78 Å. U.; V, 6132.06 Å. U. The spectrum emitted by the LaO mol. is exceedingly complex; more than 300 bands have been observed. The analysis of LaO bands agrees in the main with that given earlier by JEVONS (*C. A.* 23, 5414). W. F. MCGEE

Theoretical value of the fundamental vibration quantum for gaseous alkali iodides. H. J. VAN LESTWIJ, *Z. Physik* 66, 241-3 (1930).—A math. calcn. from the standpoint of wave mechanics, on the assumption that the carbon is small with respect to the anion, and neglecting polarization effects. The fundamental frequency ν , 10^{-13} in sec^{-1} is for LiI 9.59, NaI 4.99, KI 3.53 and RbI 2.57, which gives λ ($2\pi\nu$) as 509, 265, 187 and 136, resp. These agree with expt. only in the case of the first 2. G. M. E.

Rotation oscillation spectrum of acetylene. I. Band analysis. K. HANSEN AND R. MEYER, *Z. Physik* 64, 131-61 (1930).—In order to study the configuration and form of oscillation of simple gaseous mols. the spectrum of C_2H_2 has been investigated. The spectrum was photographed in the range 9000-7000 Å. U. with large dispersion (2.6 Å./mm.). Three absorption bands were found at wave lengths 7657, 7656 and 8620 Å. U., of which the 1st was by far the most intense. The absorption lines were measured to 0.01 Å. U. The spectrum was analyzed. The P and R branches were very simple, no fine structure of the lines being visible within the above limit of error. There were marked intensity changes. The simple structure of the bands leads to the conclusion that the 4 atoms of the mol. are arranged in a straight line, so that the method of rotation of a diat. mol. has to be dealt with. The fact that only 1 line [$P(0)$] is missing in the spectrum is in agreement with the usually accepted electronic structure of C_2H_2 and shows the analogy of the C_2H_2 mol. $\text{HC}-\text{CH}$ to the N_2 mol. $\text{N}::\text{N}$. The moment of inertia of the mol. in the fundamental state is 23.309×10^{-40} . The nuclear distance for $\text{C}-\text{H}$ is 1.08×10^{-8} cm., in agreement with that arrived at for the CH_4

The Raman effect in certain substances with a new apparatus. B. VENKATESACHAR AND L. SIBAIYA *Indian J. Physics* 5, 747-54 (1930).—A new app. was designed to meet the following requirements: (1) It must permit the use of solids; (2) it must use a strong ultra-violet line for excitation; (3) it must allow the use of elevated temps. This was accomplished by enclosing the specimen in a capsule which is placed in a fused silica tube which in turn is fixed inside a larger Pyrex tube. The exciting arc is formed in the annular space between the tubes. The temp. is changed by blowing heated air through the inner tube. With increasing temp. the lines become more diffuse and the intensity of the anti-Stokes lines is enhanced as is also that of the anti-Stokes side of the wings accompanying the unmodified lines. The spectra of calcite and aragonite are given. Halite and fluorite gave neg. results. C_2H_2 gas gives lines corresponding to the infra red absorption bands at 2.45μ and 7.5μ . The band at 13.7μ has no corresponding Raman line. J. B. AUSTIN

The Raman spectrum of nitric oxide. F. RASETTI *Z. Physik* 66, 646-9 (1930); cf. *C. A.* 24, 5627.—NO appears to be the only chemically stable diat. mol. with a $^2\Pi$ term as its ground state. The Raman spectrum of NO was observed and shows a pair of lines sep'd by the same $\Delta\nu$ as are the components $^2\Pi_{1/2}$ and $^2\Pi_{3/2}$, as det'd by observations on the absorption spectrum of NO. The Raman lines corresponding to the rotation spectrum of NO were also observed, and although only partially resolved, are in qual. agreement with those required by theory. C. C. KIESS

Raman effect in solutions of sodium nitrate of varying concentrations. V. STERLING AND E. R. LAIRD *Phys. Rev.* 36, 148-9 (1930).—The Raman spectrum of 5, 10, 30% sat'd solns. of NaNO_3 shows a line with a frequency difference of $1049 \pm 6 \text{ cm}^{-1}$. It is attributed to the NO_3 ion. BERNARD LEWIS

Studies in Raman effect. X. The Raman spectra of organic substances. A. DADIEV AND K. W. F. KOHLRAUSCH *Sitzb. Akad. Wiss. Wien, Aht. IIa*, 139, 459-72 (1930), *Monatsh.* 57, 225-40 (1931), cf. *C. A.* 25, 875.—Measurements of propyl-, diethyl-, diphenyl-, trimethyl- and triethylamines, aniline, formamide, dicyanodiamide, methylnitroamine, chloral, ethyl chloroformate, chloroacetyl chloride, dimethyl- and diphenyl ethers. The spectra of 16 amines are compared. L. ONSAGER

Raman effect in methyl halides. GAJENDRA NATH BALL *Z. Physik* 66, 257-60 (1930).—The 2 bands at 16.8 and 19.6μ for CH_3Br and CH_3I which were beyond the range of Bennett and Meyer (cf. *C. A.* 23, 2657) but which were predicted by them, have been found. Complete measurements to 0.001 mm. have been made of Raman spectra for all of the methyl halides, and are compared tabularly with B. and M. and Kohlrausch. GREGG M. EVANS

Raman effect in trimethylethylene. D. FRANKLIN AND E. R. LAIRD *Phys. Rev.* 36, 147 (1930).—Raman lines in trimethylethylene were found which correspond to infra red lines at $3.44 \pm 0.05\mu$ and $8.4 \pm 0.2\mu$. An anti-Stokes line was found at 8.4μ . BERNARD LEWIS

Raman spectra of organic sulfides. V. N. THATTE AND A. S. GANESAN. *Nature* 127, 306 (1931).—The Raman spectra det'd are as follows, expressed in cm^{-1} : Et_2S 652, 1061, 1282, 1439, 2923; allyl sulfide, 410, 588, 741, 917, 1011, 1101, 1210, 1291, 1312, 1420, 1534, 1636, 3007, 3088. The frequency 652 seems characteristic of sat'd. sulfides; the frequency 751 a characteristic of unsat'd. sulfides. A. F.

Color and constitution from the standpoint of recent electronic theory. IV. Various types of anomalous nitration; chelation; complex salt formation; unusual stability of substituted aminophenols; direct diacetylation of amines; inhibited hydrolysis of sulfonic acids; reaction of caustic potash with benzaldehyde. H. H. HODGSON. *J. Soc. Dyers Colourists* 46, 183-8 (1930), cf. *C. A.* 24, 4223.—The reaction of HNO_3 with dimethyl *p*-toluidine to give monomethyl *p*-toluidine-*N*-nitroscamine and 3 nitro-dimethyl *p*-toluidine is described and explained. It was observed in a series of isomeric and homologous comp'ds. that the picrates in which the greatest salt formation occurred had the higher m. ps., while those where the union was of a more chelate type had lower m. ps. The anomalous behavior of the picrate of dimethyl *p*-toluidine is explained. The direction of nitration of anilides and phenols, *o* or *p*, is det'd by the strength of the cation of the nitrating agent used. Strong cationoid reagents cause para nitration and weak ones ortho. The stability of the 3-halo-2(4, and 6)-aminophenols is considered to be due to the prevention of inner salt formation by the constraint imposed on the electrons of the N by the halogens. The fact that acetylation of 2,6-dichloro-4-nitroaniline produces always the di- instead of the mono-acetylated product is explained by assuming a mol. rearrangement after the introduction of the first acetyl group which greatly facilitates further acetylation. The effect of other ring substituents on the ease of hydrolysis of sulfonic acids is explained. The reaction of caustic potash with

BzH is explained by assuming the formation of an isomeride of benzoin which subsequently decomposes to form benzyl alc. and BzOH. T. G. HAWLEY

Chemical decomposition by radiation. G. HARKER. *J. Cancer Research Comm. Univ. Sidney* 2, 111-8, 160-81 (1930).—A review with bibliography. Radiation as applied in therapeutics is also discussed. G. SCHWACH

Mechanism of the chemical action of x-rays in aqueous solutions. O. RISSÉ. *Strahlentherap.* 34, 578-81 (1929). cf. *C. A.* 24, 5602.—Ultra violet irradiation of water whether in presence or absence of O, did not lead to the formation of H_2O_2 unless Zn oxide was present, α or β irradiation of pure water leads, however, to its formation. H_2O_2 is decomposed by all 3 radiations. The reaction depends on the dissolved O; hence the presence of dissolved O in irradiated solns. may cause changes dependent on the primary formation of H_2O_2 . The conversion of ferrous into ferric salts is an example. The mechanism of the formation of H_2O_2 is discussed. B. C. A.

A new method of preparation of phosphorescent zinc sulfide. JOSEF ERNIC. *Chem.-Ztg.* 55, 31 (1931).—After illumination, ZnS becomes phosphorescent if it be otherwise as nearly pure as possible, but mixed with a small quantity of a heavy metal, particularly Cu, and heated to above 900° to convert the ZnS to cryst. form. Very nearly pure ZnS contg. less than 1/100,000 per 1000 parts of heavy metals does not phosphoresce. The amt. of Cu required to make ZnS phosphoresce is 0.00001-0.001 part, outside these limits there is no phosphorescence. The new process consists in purification of $ZnSO_4$ soln. by causing a slight pptn. by addn. of H_2S or Na_2S , then heating to ppt. all metal sulfides less sol. than ZnS. Sulfides more sol. than ZnS remain in soln. and can only be removed later, since they form during the reaction. ZnS so formed does not phosphoresce, but can be made to do so by addn. of the required amt. of Cu. The product is superior in luminosity to the com. product. E. M. SYMMES

Extinction of fluorescence of solutions of dyes by electrolytes. A. V. BANOV. *Z. Physik* 64, 121-34 (1930). cf. *C. A.* 24, 2052.—The application of the theory of Vavilov concerning collisions of the 2nd kind to the extinction of fluorescence in solns. of dyes by adding electrolytes does not agree fully with exptl. results. This leads to the search for some other cause of extinction which is to be found in the "salting-out" effect of the electrolyte on the dye. Exptl. work supports this view. Calcula. of the sum of the radii of the mol. of the dye and of the ion of the electrolyte give results in agreement with those arrived at by other methods. The extinction curves reveal a connection between the concn. of the electrolyte, c , and the intensity ratio after and before adding the electrolyte, which can be expressed in the form $L/L_0 = e^{-k\sqrt{c}}$. B. C. A.

Photoreaction of hydrogen and iodine monochloride. S. E. ASHLEY AND WILLIAM WEST. *Nature* 127, 308 (1931).—Impurities reduce the rate to a small but measurable value, confirming Mellor and Iredale (cf. *C. A.* 25, 1736) but not agreeing with Rolfe and Lindquist (cf. *C. A.* 24, 4224; 25, 1736). The photoreaction between CH_4 and ICl is very rapid. ARTHUR FLEISCHER

Photochemical kinetics of chlorine detonating gas. Oxygen-free gases. MAX BODENSTEIN AND WALTER UNGER. *Z. physik. Chem., Abt. B.* 11, 253-78 (1930); cf. *C. A.* 8, 452, 20, 3646, 21, 3527.—H was prepd. O-free by passage over heated Pd asbestos and diffusion through Pd. Cl was liquefied and swept out with pure H. Fractionation under low pressure was also used. The reaction between H and Cl was followed by the methods of Dux and Thon in a quartz vessel, illuminated with light of 4360 Å. Without O the reaction followed the mass law, $d[HI]/dt = k[H_2][Cl_2]$ and considering the intensity, of the absorbed light $d[HI]/dt = kI[H_2]$. The breaking of reaction chains by reactions on the vessel walls was found proportional to $[Cl]$. To account for this the hypothesis of diffusion of Cl atoms to the walls and recombination there appeared untenable. This effect was due to formation of a volatile compd. of Cl and Si from the container walls. The influence of temp. was studied from 30° to 167° by a streaming method. The effect of small amts. of O was further investigated. L. P. HALL

Cystine and its absorption spectra. D. BAÑER. *Anales soc. españ. fis. quim.* 28, 1386-1406 (1930).—A spectiographic study of the absorption spectra of cystine and alanine in HCl soln. was made by a sector photometer with a small quartz spectrograph to analyze the light absorbed by the ammo acid in a Baly tube. The light source was a discharge through pure H₂ at a pressure to give a continuous spectrum. The results are shown graphically by plotting the mol. absorption coeff. against wave length. They agree with Ward (*C. A.* 18, 1117) within exptl. error. Curves obtained with cystine and alanine are similar, and show no selective absorption. The alanine absorption begins further in the ultra violet, and the curve is incomplete because of limited instrument sensitivity. Intensity of absorption increases with short wave lengths.

from the detn of wave length at which absorption begins, energy of dissociation $h\nu = 4.9 \times 10^{-13}$ ergs per mol., and the heat of dissociation = 70,800 cal per g-mol., and for alanine $h\nu = 5.75 \times 10^{-13}$ and $H = 83,000$, resp. The structure of cystine is aliphatic. Cystine is more easily dissociated than alanine. E. M. SYMMES

Absorption spectra of bibenzyl and its derivatives. MICHELE RAMART-LUCAS AND J. HOCQUET *Compt rend* 192, 53-5(1931)—Bibenzyl and its derivs., prepd. by the action of Mg or Na on $BzCl$ or similar compds gave compds with 2 absorption bands, (a) between 3300 and 2700 Å U and (b) between 2700 and 2300 Å U. Similar compds prepd by the action of $PhCH_2CH_2Br$ on benzene in the presence of $AlCl_3$ gave compds with a band at 2700-2300 Å U only. Further investigation showed band a to be due to stilbene. WALLACE R. BRODE

The action of radiation on colloids. III. Action of ultra-violet light on emulsions. P. C. SENNA AND B. G. GANGULI *Kolloid-Z* 54, 147-50(1931) cf. *C. A.* 24, 540, 5632—Emulsions of $C_{11}H_{16}$, benzene and CS_2 in water with Na oleate as the emulsifier sepd. into 2 layers on illumination. The η of the illuminated emulsion increased, indicating that the breaking is due to a partial destruction of the emulsifying agent. A. F.

Colloidal nature of the coloring substance in colored rock salt. M. SAVOSTYANOVA. *Z. Physik* 64, 262-78(1930)—Mie's theory was applied to the system Na-NaCl, and the absorption and diffusion spectra for different particle sizes were calcd. and compared with the exptl. curves obtained for colored rock salt. The absorption spectra of the following kinds of colored salts were detd. by the spectrophotometer method. (a) natural blue salt, (b) additive (in liquid Na) colored salt and (c) salt obtained from yellow salt by the action of heat and pressure under the influence of light. The exptl. data agree with the assumption that the colors are due to particles of colloidal Na. Besides the absorption maxima due to the presence of colloidal Na, a further max. was found for yellow salt in the infra red, the explanation of which is not clear. B. C. A.

The influence of light on the flocculation of colloidal solutions in fluorescent media. Role of autoxygens. AUGUSTIN BOUTARIC AND JEAN BOUCHARD *Compt rend* 192, 357-8(1931), cf. *C. A.* 1738—In continuation of the work previously described, the parallelism between the effect of some electrolytes in depressing the fluorescence of fluorescein solns and their effect on the flocculation of As_2S_3 sols by light in presence of fluorescein is studied. The difference Δt between the time required for flocculation in the dark and in daylight or ultra violet is approx. proportional to the logarithm of the fluorescing power of the mixt. of fluorescein and the electrolyte. H_2SO_4 gives the greatest effect, LiCl none. Tannin, hydroquinone, phenol and cresol in presence of LiCl also reduce both fluorescing power and Δt . Fluorescing power was detd. by the fluorometer of Perrin. The possible biof. significance is discussed. K. V. T.

Luminescence analysis (BEUTEL, KUTIELNIGG) 26. The molecular structure of triatomic gases (MAHANTI) 2. The propagation of flame in electric fields (GUÉVAULT, WHEELER) 2. Literature on the use of the x-ray (POLANSKY) 9. Separation of salts of Ra and Ba (Ger. pat. 515,681) 18.

FAJANS, KASIMIR. Radioelements and Isotopes: Chemical Forces and Optical Properties of Substances. George Fisher Baker Non-Resident Lectureships in Chemistry at Cornell University Series. New York. McGraw-Hill Book Co., Inc. 125 pp. \$2.50

FLINT, H. T. Wave Mechanics. 2nd ed., revised and enlarged. London: Methuen. 3s. 6d., net.

JOYE, PAUL. La matière et l'atome. Fribourg: l'Oeuvre de Saint-Paul. 19 pp.
WEYL, HERMANN. Gruppentheorie und Quantenmechanik. 2nd ed. Leipzig: S. Hirzel. 338 pp. Bound, M. 26. Reviewed in *Phys. Rev.* 37, 783(1931)

Concentrating radioactive substances. KURT WEIL and KURT PETERS. *Ger.* 518,205, Dec. 21, 1929. The concn. of active substance in radioactive carbonate mixts. is raised by fractional thermal decompn. followed by washing with water. Thus, a mixt. of $RaCO_3$ and $BaCO_3$ may be heated to 400° in *vacuo*, cooled and washed with hot water, a residue richer in Ra than the initial mixt. being obtained.

Activating metals or alloys. OSKAR RIED. *Austrian* 121,243, Sept. 15, 1930. Metals or alloys are subjected to radiations of suitable wave length, e. g., to ultra-violet radiations, whereby they are physically and biologically activated. Numerous applications of the invention are mentioned. Thus, the growth of yeast may be influenced by effecting the cultivation in irradiated metal vessels, or the stability of milk under trans-

port may be improved by using irradiated metal cans, or butter or cheese may be packed in irradiated metal foil

4—ELECTROCHEMISTRY

COLIN C. FINE

The institute for technical electrochemistry in the Berlin technische Hochschule
K. ARNDT *Z. Elektrochem.* 37, 218-21 (1931) G. G.

The present development of the electrochemical industry and electrometallurgy
in Italy G. MORSELLI AND O. SCARPA *Trans. 2nd World Power Conference (Berlin)* 1,
299-305 (1930) E. H.

The application of electric heat in industry. H. NATHUSIUS, *et al.* *Trans. 2nd
World Power Conference (Berlin)* 1, 390-428 (1930) E. H.

Theoretical consideration of the oxidation of pig iron during its transformation into
steel MUKIO MUKIYAMA *Tetsu to Hagane (J. Iron and Steel Inst. Japan)* 16,
1041-57 (1930).—From the mass action law, M. deduced the formula for the equilibrium
const. $K' = dx/dt(1/(a + M - x)(b - x))$, where x is the density of C at time
 t , a and b are the initial contents of FeO and C, resp., and M is the diffusivity of FeO
when no reaction occurs. To eliminate M , he obtained an approx. value of $K' =$
 $1/t \log b/(b - x)$. Expts. were undertaken to obtain values for b . After heating the
mixture of pig and hematite in a resistance furnace and an induction furnace, the products
were examined by chem. analysis and microscopic observation. Results showed that
 $K' = 0.000$ at 1450° and 0.006 for 1500° in the induction furnace and 0.000 and 0.007 ,
resp., for the resistance furnace. M. KURODA

High-frequency induction furnaces for making tool steel. J. A. SULLOCH *Metal
Progress* 18, No. 4, 40-4 (1930).—A review W. A. MUDGE

A new cyanide furnace. R. F. JAMES AND GLENN COLEY. *J. Franklin Inst.* 211,
327-34 (1931).—A new elec. resistance furnace has been constructed on the hypothesis
that the maintenance of an oxidizing atm. in the heating element chamber at all times
is essential to the preservation of the elements. A drawing is shown. The test pot
had an input of 43 kw. After a service of 1900 work hrs., a check of the element showed
a power input of 42.5 kw., which is a considerable improvement over the power-consumption
curves from previous designs. No scale was found on the heating elements, but
the pot showed traces of leakage and a heavy deposit of what appeared to be cyanates
on the floor of the element chamber. Elec. tests and phys. exams. showed no deteriora-
tion of the elements. This furnace consistently produced over 4 lb. of work per kw.
hr. without any covering of the pot. A full charge of work was delivered in 18 to 20
min., which is well within the production speed desired. Appreciable quantities were
lacking of the black sludge usually formed on the surface of molten NaCN. H. C. P.

Bright anneal in artificial atmospheres. J. C. WOODSON *Elec. J.* 28, 246-7
(1931).—Electrically heated furnaces are very successfully used for bright anneals.
Partially cracked and purified coke-oven gas and treated blast furnace gas are not
entirely satisfactory for the bright anneal of certain metals. H₂, dissociated NH₃, natural
or city gas (with or without water vapor), MeOH and butane are similarly used. Blast-
furnace and coke-oven gas are used on low C, cold rolled steel, cracked butane is used
on high C steel or roller bearing stock. The artificial furnace atm. will probably soon
be adopted in fabricating and mfg. processes. W. H. BOYNTON

The electrolytic industry of Germany. V. ENGELHARDT, *et al.* *Trans. 2nd World
Power Conference (Berlin)* 1, 317-48 (1930) E. H.

Operating data on electrolytic metal production and refining. C. L. MANTELL.
Eng. Mining J. 131, 256-9 (1931).—Operating data from the electrolytic plants of the
U. S. and Canada show the compn. of electrolytes, equipment, methods of applying
the current and efficiency obtained from the power, raw materials used, and by products
produced. The tables cover the electrolytic refining of Cu, the electrolysis of Cu-
bearing solns. obtained by leaching processes, and the production of Zn, Cd, Pb, Ni, Sn, Ag
and Au. The high power cost of processes using insol. anodes is indicated in the tables.
The diaphragm used in Na₂ production causes an increase in cell voltage with attendant
high power costs. W. H. BOYNTON

Factors affecting the smoothness and mechanical properties of electrodeposited
copper. A. W. HOTHERSALL. *J. Electroplaters' and Depositions' Tech. Soc.* 6, 95-122,
Metal Ind. (London) 38, 343-6 (1931).—Growth formation, unequal current distribu-
tion, maintenance of an adequate supply of Cu ions at the cathode, factors which
tend to reduce the rate of Cu deposition in practice and the detn. of the influence

of the more favorable depositing conditions on the mech. properties of the deposits are discussed. From solns. contg 200 g $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ and 50 g H_2SO_4 per l. (Soln. A) (temp. = 26°), variations in c. d. from 15 to 200 amp/sq ft, with or without air agitation, caused little change in the Brinell hardness (from 42 to 52) of the deposited Cu. Increasing c. d. and increasing acid concn. caused slightly harder deposits to be given. A hard deposit (Brinell no. 81) was obtained at 15 amp/sq ft. from a soln. contg 200 g $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ and 20 g $(\text{NH}_4)_2\text{SO}_4$ per l. and no free acid. A very hard, smooth and exceedingly brittle deposit of polished appearance was obtained at 16° and 15 amp/sq ft. from a neutral soln. contg CuSO_4 and NH_4NO_3 . The unequal current distribution (particularly marked in Cu deposition because of the steepness of the cathode potential-c. d. curve), together with the formation of nodular growths, due to (a) irregular crystal growth of the deposit and (b) adherence of electrically conducting particles (formed in the soln.) to the cathode surface, and the effect of agitation on the above factors, is discussed in detail. Precautions (e. g., warm solns., filtration, adjustment of agitation to c. d. being used and correct pickling of cathode basis metal) which should be taken in solns. contg. no addn. agents in order to secure smooth Cu deposits are reviewed. The effects of various addn. agents in Cu deposition, the control of such substances, the mechanism of their action and the mech. properties of Cu deposited from solns. contg. addn. agents are discussed. Addns. of $\text{Al}_2(\text{SO}_4)_3$ up to 32 g/l. to Soln. A (above) had no effect on the mech. properties of deposits obtained on a rotating cathode (peripheral speed 450 ft./min.) at 50 amp/sq ft., nor on the cathode potential. The ductility of the Cu deposited from Soln. A upon the rotating cathode was considerably increased by the addn. of 0.06 g/l. Cl (as NaCl), the % elongation increasing from 33 to 48, and at the same time the tendency for finely divided Cu to form in soln. was retarded. Although beneficial in agitated baths, in still baths the Cl induced the growth of thin needle-like single crystals. The addn. of the requisite amt. of Ag_2SO_4 soln. to remove the Cl from such baths caused them to yield satisfactory deposits subsequently. Addns. of gelatin and peptone to Soln. A caused the deposits to have smaller crystals and to be smoother, harder and more ductile. The Brinell hardness of the deposits from still baths increased from 42 to 83–93 by 2 p.p.m. of these substances, and 0.01 g/l. gelatin raised the hardness to 123. Dextrin, up to 10 g/l., had little influence on the hardness and crystal size of the Cu from freshly prepd. solns., but when "aged," such solns. gave harder and more brittle deposits. Cu from freshly prepd. solns. (A) contg. up to 10 g/l. phenol (as phenolsulfonic acid) was similar in properties to, but much smoother than, that from solns. contg. no addn. agents. On continued electrolysis the effect of the phenol and also of the gelatin and of peptone disappeared. The hard deposits given in "aged" solns. contg. phenol were more ductile than those formed without addn. agents. Dextrin is concluded to be the least suitable addn. agent; the pronounced effects of gelatin and peptone make it necessary for very small quantities to be used—with the consequent difficulty of estn. and control of the amt. present—whereas phenol can be used up to 10 g/l. and can be detd. analytically (Hull and Blum, *C. A.* 25, 460). Tables are given showing the hardness of the deposits from solns. free from addn. agents, and also the effect of addn. agents on the hardness and tensile properties of the deposits from still and from air-agitated solns. Cathode potential-c. d. curves for solns. with and without addn. agents, both for freshly prepd. and for "aged" solns., are given, as well as photographs of some deposits (including nodules and single crystals) and their micro-structures. A bibliography is appended. EDWARD B. SANIGAR

The reduction of arsenic acid and arsenates to arsine at the mercury cathode. I. Wm V. LLOYD *Trans. Faraday Soc.* 27, 89–94 (1931).—A rotary commutator was used to connect the cathode to the source of current and a potentiometer alternately 3000 times a min. The overpotential decreased by 0.5 v. to that of As, although analysis of the H_2 cathode showed no As present. The decrease in overpotential was less when an uninterrupted current was used, and varied directly with increasing concn. of H_2AsO_4 . The quantity of AsH_3 produced increased with increasing H_2AsO_4 concn. The efficiency of the reduction process, based upon the H converted into AsH_3 at the cathode, varied from 1.71 to 14.1%. *Cf. C. A.* 24, 1582. CURTIS L. WILSON

Temperature measurements on working electrodes. IV. B. BRUÏS. *Z. physik. Chem., Abt. A*, 153, 309–20 (1931). *cf. C. A.* 24, 3179.—The electrolytic Peltier effect which occurs in working electrodes was evaluated with the help of a dynamic calorimetric method, a detailed discussion of the theory of the method being included. The Peltier effect is measured for $\text{Hg}|\text{Hg}_2^{++}$ in perchlorate and nitrate solns. at different concns., and the results are applied to the calcn. of the partial entropy of Hg_2^{++} ion for the same concns. For solns. of unit activity this is 29 ± 1 cal/degree. H. STOERTZ

The composition of manganese amalgam and manganese-silver alloys in relation to

the electrode potential of manganese. HARVEY D. ROYCE AND LOUIS KAULFNBURG *Trans. Electrochem. Soc.* 59 (preprint) 23 pp (1931).—Four Mn-Ag alloys were prepared from highly purified materials. Electrode potentials and replacing power of these alloys have been measured. The replacing power and potential of alloy D (31 at. % Mn) were correlated with alloy structure by means of photomicrographs and thermal diagrams. None of the Mn-Ag alloys yields electrode potential approaching that of Mn amalgam. Mn amalgam consists of solid $Mn_{11}Hg_2$ in equilibrium with a very dilute solution of Mn in Hg (0.0032% Mn). The behavior of Mn amalgam in contact with Mn salt solutions has been interpreted on the basis of these figures. A new compound of Mn and Hg isolated from the dilute amalgam by the application of pressures exceeding 150,000 lb per sq in (10,500 kg/cm²) and having the formula $Mn_{11}Hg_2$ is described. $Mn_{11}Hg_2$ constitutes the stable solid phase in the system Mn-Hg at temperatures between 80° and 100°. Pressure-concentration curves are presented to justify the assumption that the solid phase obtained by pressure is $Mn_{11}Hg_2$ contaminated with a small percentage of the liquid amalgam. C. G. F.

Bright silver plating from the cyanide bath. A preliminary note. L. C. PAN. *Trans. Electrochem. Soc.* 59 (preprint) 5 pp (1931).—A number of chemicals were investigated with regard to their brightening effect upon electrodeposits of Ag from the cyanide bath. NH_4OH was found to have a decided whitening effect upon silver deposits, and was especially effective in an electrolyte containing large amounts of impurities. Na thiosulfate at a concentration of about 1 g/l was found to be a good brightener for the Ag bath, and more effective than CS_2 . The combination of NH_4OH and Na thiosulfate in a Ag bath produces a bright deposit similar to bright Cr, even in the presence of impurities which ordinarily cause dark yellow deposits. With the above brighteners, bright deposits are possible at moderately high current densities. C. G. F.

The change in reflecting power caused by tarnishing electrodeposited silver-cadmium alloys. LAWRENCE F. STOLT AND WILLIAM G. TRUMWEL. *Trans. Electrochem. Soc.* 59 (preprint) 13 pp (1931).—Ag-Cd alloys can be deposited from a commercial Cd plating bath to which AgCN has been added. The composition of the alloy depends on the Ag concentration in the bath (if the Cd content remains constant) and on the cathode current density. The co-deposition of Ag lowers the reflecting power of electrodeposited Cd. The tarnishing of the alloys depends both on the composition and on the rate of deposition and is materially less than that of pure Ag. A Ag-Cd alloy containing about 24% of Ag recommends itself as the best general purpose reflector. Its reflecting power is equal to that of pure Cd, and it is considerably harder and more resistant to abrasion than pure electrodeposited Cd. C. G. F.

Progress of the hot nickel solution. OLIVER P. WATTS. *Trans. Electrochem. Soc.* 59 (preprint) 5 pp (1931).—A report is made of the volt and operating conditions of many of the hot Ni plating solutions now in use in the U. S. and Canada. Since there was a period of 5 years between the announcement of the advantages of heating the Ni plating solution and its commercial adoption, the reported growth to 158,000 gallons has occurred in 10 years. Opinions of platers on the respective merits of the hot and cold Ni solutions are included. C. G. F.

Control of acidity in low-pH nickel-plating baths. L. C. PAN. *Trans. Electrochem. Soc.* 59 (preprint) 8 pp (1931).—For the control of acidity in Ni plating baths of low pH, the titration method, including a "control chart" for direct readings, was found to be most practical in a commercial plant. It is based upon the assumption that for a given bath, whose composition is kept approximately constant, the pH value will be the same as long as the acidity is maintained at the same point. The "control chart" also provides a direct method of regulating and correcting the acidity or pH, eliminating all labor of computation and all guesswork. C. G. F.

Chromium deposits directly on aluminum. HAROLD K. WORK AND CHARLES J. STUNDER. *Trans. Electrochem. Soc.* 59 (preprint) 7 pp (1931).—Thin layers of Cr which have fairly good resistance to salt spray and atmospheric corrosion may be electrodeposited directly on Al. Such deposits have a gray color on leaving the plating bath and must be buffed if a higher luster is required. These deposits protect Al to a marked degree from alkaline corrosion. Heavy deposits for abrasion resistance have also been applied to certain Al alloys. C. G. F.

The electrodeposition of chromium and the influence of the cathode metal. J. W. CUTHBERTSON. *Trans. Electrochem. Soc.* 59 (preprint) 24 pp (1931). cf. C. A. 25, 35.—In Cr plating the best results and greatest flexibility and reliability in operation are obtained from a solution of CrO_3 and H_2SO_4 . The molarity of the CrO_3 should not be greater than 3.25, and the ratio of this to SO_4 normality should be about 58 to 60. The tank should be lined with hard Pb, the lining not being anodic to the solution except as a

temporary measure. Pb is the only material that can be considered for anodes. The temp. should be reasonably high, but not above 45° , as then the Cr^{III} equil. conc. reaches too high a figure. Care should be paid to the ratio between anode and cathode surfaces. The anode area should be 1.5 to 3 times that of the cathode, and anodes should be disposed with respect to the cathode size and shape. Even under the best of conditions, the cathode current efficiency is not above 16%. The nature of the cathode is of the greatest moment. It should be of a metal of high H overvoltage corresponding to that of Cr. Ni is the material mostly used now, but it is far from ideal. For many purposes Cu is a far better undercoating. It is impossible to increase materially the throwing power of CrO_3 solns. The best way to improve upon existing conditions is to increase the plating range in which bright plate is obtainable. Except in the case of Cu little success has been obtained from Cr deposits direct on the base metal. This may be due to the inherent porosity of these deposits. Direct deposits can be improved by heat treatment, but this does not lend itself readily to commercial practice. C. G. F.

Electrodeposition of chromium-iron alloy. G. FUSEYA AND K. SASAKI. *Trans Electrochem. Soc.* 59 (preprint) 24 pp (1931).—See C. A. 25, 1164 (1931). C. G. F.

Chromium-plated molds. H. E. SROCKFISCH. *Am. Machinist* 73, 553-5 (1930).—The advantages are given of Cr-plated surfaces in the molding of glass brick and mica to reduce wear and improve finish, and the compn. and treatment of molds are discussed. E. I. S.

The electrodeposition of tungsten from aqueous solutions. COLIN G. FINE AND FRANK L. JONES. *Trans Electrochem. Soc.* 59 (preprint) 23 pp (1931).—The W deposited from aq. solns. is smooth, hard and coherent and has a high luster. Like Cr, it needs no polishing if the plated article was previously polished. W has remarkable acid-resisting properties, which make it desirable as a protective coating for other metals. Several types of solns. have been investigated in regard to their usefulness as W plating baths. Solns. of W salts in org. solvents are unsuitable. Aq. solns. of an acid nature do not give satisfactory metal deposits. Alk. solns. contg. alkali tungstates are preferably used as plating baths. The temp., cathode c. d. and H-ion concn. of the soln. must be controlled. Alloys of W and Fe, and of W and Ni, were electrodeposited from similar solns. contg. very small concns. of the ions of these second metals. The usefulness of highly alk. solns. as plating baths for the electrodeposition of other metals is pointed out. C. G. F.

The production of hydrogen at certain metal surfaces in relation to overvoltage. G. RAYMOND HOON AND FRANCIS C. KRAUSKOPF. *J. Phys. Chem.* 35, 786-95 (1931).—The efficiency of electrolytic reduction of KClO_4 (in 0.493 N soln.) was detd. concurrently with the cathode potential V (measured against a satd. calomel electrode with a KCl bridge), various metal cathodes and several c. ds. being used. The data for different cathode materials conform to the same relation between V and reduction efficiency; the efficiency decreases with increasing V , and is negligible for $V > 1.5$ v. Measurements of the decompn. potentials of H_2SO_4 (0.1 N), KCl (0.493 N) and KClO_4 (0.493 N) with Pt anodes and various cathodes are given and compared with the reduction expts. L. O.

Electrolytic production of hydroxides and hydrogen. R. SAXON. *Chem. News* 142, 85, 113-4, 149-50 (1931).—These articles are statements of the phenomena occurring when pure H_2O (contg. no acid or salt for conduction) is electrolyzed, 240 v. being used, with Pt or C cathodes and anodes of various metals. The gas liberated at the cathode was collected over H_2O and found to be H only. The solns. took on various colors depending on the metal used as anode, and metal hydrides were formed in the solns. Expts. with the following metals as anodes are reported. Fe.—The soln. turned amber, then green; green $\text{Fe}(\text{OH})_2$ formed in the soln., while a scum of $\text{Fe}(\text{OH})_3$ floated on the soln., even that in contact with the cathodic H . The soln. was acid to methyl orange but gave no test for Fe unless acid was added. It is concluded that the natural tendency is for $\text{Fe}(\text{OH})_2$ to be formed, but that the nascent (not the molecular) H reduces this to $\text{Fe}(\text{OH})_3$. Cu.—The soln. turned green and, later, green-blue to purple. The nascent H reduced the $\text{Cu}(\text{OH})_2$ first formed to CuOH and then to Cu which floated on the surface of the soln. The $\text{Cu}(\text{OH})_2$ first formed was found to be electropositive. Co.—The soln. became pale bluish, a green-blue scum, thought to be $\text{Co}(\text{OH})_2$, collected on the H_2O in the tube over the cathode. Red-brown flakes of $\text{Co}(\text{OH})_3$ appeared in the soln. followed by brown and black masses of Co_2O_3 and CoO . Ni.—An apple-green film of $\text{Ni}(\text{OH})_2$ was observed followed by a green scum of $\text{Ni}(\text{OH})_3$ and, later, floating specks of Ni_2O_3 , followed quickly by green NiO . From the expts. it is concluded that the reductions are brought about only by nascent H (not by molecular H), and that Ni is preferentially divalent, Fe preferentially tri valent, while Co is bi- and ter-valent without preference. EDWARD B. SANGAR

The use of antimonial lead for accumulator grids: a cause of self-discharge of the negative plates. J. T. CRENNELL AND A. G. MILLIGAN. *Trans. Faraday Soc.* 27, 103-12 (1931).—It was observed that the neg. plates of Pb accumulators at high temps. and high acid concns. lost capacity as the cells became older. The use of exptl. cells with pure Pb grids and with grids of antimonial Pb indicated that the "aging" of the neg. plate was due to the soln. of Sb from the alloy of the pos. plate and its deposition on the neg. plate. The low H₂-overvoltage at an Sb surface reduced the current efficiency of the plate by allowing evolution of H₂. Sb added to the electrolyte produced the same effects. No quant. relation between the total Sb added and the rate of sulfation of the neg. plate could be established. It is concluded that antimonial Pb should not be used for the pos. grids of Pb accumulators. CURTIS L. WILSON

Diatomite as a filler in battery boxes. PAUL HATMAKER. *Bur. Mines, Rept. of Investigations* 3078, 2 pp (1931).—Diatomite for battery fillers must have < 8-10% acid-sol. material and < 0.75% Fe and be low in clay and org. substances. Mn and MgO are objectionable. Present consumption for this use is 10,000-15,000 tons annually. ALDEN H. EMERY

Electricity in modern dairy plants. A. J. DREUX AND H. C. BRUNER. *Elec. Eng.* 50, 255-7 (1931).—The elec. current passes through the milk in the Electropure method of pasteurizing. The application at the Rock McJunkin Dairy in Pittsburgh, Pa., is described, also other uses of electricity in such plants. W. H. BOYNTON

Electrochemistry applied to glass. ALEXANDER SILVERMAN. *Trans. Electrochem. Soc.* 59 (preprint) 9 pp (1931).—Accessory applications of electrochemistry to glass manif. are discussed e. g., cond. as a control of viscosity, elec. pptn. in purifying producer gas, Cr plating of molds, and electroplating of the finished product with various metallic coatings. A review is given of studies that have been conducted on the electrochem. properties of glass, such as conductance through the glass in d. c. electrolysis. C. G. F.

Influence of traces of hydrogen on sparking potential of helium. A. GÜTHER-SCHULZE AND F. KELLER. *Z. Physik* 65, 219-23 (1930).—The disappearance of the last traces of H in glow lamp rectifiers caused an increase in the flash-over voltage from 182 to 330 v. This is explained by reference to the four fold longer mean free path of the H ion over other positive ions. GREGG M. EVANS

Current collection in hydrogen atmosphere. R. M. BAKER. *Elec. Eng.* 50, 266-7 (1931).—Test runs of commutators and slip rings operating in H₂ gas reveal interesting characteristics and lead to the following conclusions. A commutator machine designed and adjusted for good commutation in air will operate satisfactorily and give good brush life in H₂. If a commutator machine must spark in H₂, the brush life may be increased many times by keeping the relative humidity below 10%. The contact resistance of C brushes on a commutator may be lowered by the introduction of H₂. C or graphite brushes cannot be operated satisfactorily on plain tool-steel slip-rings running in H₂. The contact drop between a C or graphite brush and a brass slip-ring may be 10 times as high when the ring is running in air as when in H₂. W. H. B.

The development of an electron-emitting alloy. O. S. DUFFENDACK, R. A. WOLFE AND D. W. RANDOLPH. *Trans. Electrochem. Soc.* 59 (preprint) 17 pp (1931).—It was discovered that irregularities in the sparking voltage of spark gaps were due to non-uniformities in the compn. of the electrode materials. Small inclusions of Mg in the Ni alloy ordinarily employed for spark plug electrodes were found to be distributed at random through the wire and to affect the sparking voltages. A thorough investigation established a complete coordination between the sparking voltage of a gap and the work function of electron extr. from the cathode surface. The sparking voltage of a standardized gap increases when the thermionic electron emission of the electrode wire, measured under standardized conditions, decreases. After many expts. a Ni alloy contg. Ba has been developed which is homogeneous in compn. and readily reproducible. Spark gaps having electrodes made of this alloy have very const. sparking voltages. The alloy has marked electron-emitting properties that make it valuable for use in various elec. devices. Its uses in several ways are being developed. C. G. F.

The decomposition of iron oxide in vacuum tubes. E. R. WAGNER. *Trans. Electrochem. Soc.* 59 (preprint) 3 pp (1931).—An oxidized Fe anode in a vacuum tube, type "245," liberated large quantities of gas during the operation of the tube. If the potential gradient is increased to 300 v., the emission of gas is almost instantaneous. The phenomenon does not appear to be a purely electrolytic one, but rather due to electron bombardment. C. G. F.

New photoelectric cells with potassium. ANON. *Ann. chim. anal. chim. appl.* 13,

72-5(1931).—In the new cells, the cathode is a deposit of K on the interior wall of the bulb, and the anode consists of straight or curved metallic filaments. Each electrode is connected with an external circuit of high resistance (100,000 ohms to 100 megohms) and to a storage cell which creates between the electrodes an elec. field to accelerate the electrons emitted from the cathode. In the dark, the cell presents infinite resistance and permits the passage of no current in the external circuit. The very feeble current produced by the action of light is amplified by a lamp with 3 electrodes, and its variations can be measured directly with a milli-ammeter. The photoelec. cell and the triode amplifier are mounted in a metallic case, and the cells are preferably filled with a gas like A. W. T. H.

Production of color from tungsten lamps. M. LUCKIESH *Elec. World* 97, 645 (1931).—In order to obtain colored light from a lamp it is necessary to filter out the undesired colors, but it is impossible to obtain light of a pure, spectral color by means of filters. W. H. BOYNTON.

The condensation of hydrocarbons by electrical discharge. A study of the M/N -ratio for methane, ethylene and acetylene. S. C. LIND AND GEORGE R. SCHULTZ *Trans. Electrochem. Soc.* 59 (preprint) 16 pp. (1931), cf. *C. A.* 25, 1440.—An accurate method is given for taking into account all temp. effects which have involved some uncertainty in judging the pressure changes during the expts. in the course of the earlier investigations. The properties and the applicability of the system employed for the measurements, namely, the ozonizer type of discharge, are discussed. For the acetylene condensation, a neg. catalytic wall effect of the solid reaction product on the condensation velocity has been discovered. The condensation of C_{11} , C_{12} , and C_{13} was studied in an elec. discharge with respect to the yield of ionization as represented in the M/N ratio of the condensation, where M equals the no. of reacting mole. and N , the no. of pairs of charges. The same proportionality was found for the M/N ratios of C_{11} and C_{12} (2/20) whereas the value of the ratio for C_{13} is exactly twice as high (10 instead of 5) as has been found in the α -ray work of Lind and co-workers. Recent investigations, which have been made since this paper was first written, indicate that the correlation between α -ray work and elec. discharge finds its best basis in the assumption that not the chem. yield per ion pair is the sole controlling factor, but the product of the yield (M/N ratio) and the mol. ionization. This brings the results for C_{11} , C_{12} , and C_{13} in agreement, whereas the results of the C_{13} runs now remain out of line, though it is believed that the low values found for C_{13} will be raised upon further purification of the gas from acetone or some other inhibitor. C. G. F.

Cracking oil by electricity. HERBERT R. ROWLAND *Elec. Eng.* 50, 288 D0 (1931).—Low-grade oils formerly wasted are now "cracked" to form gasoline and other highly volatile oils. Some of the problems involved are discussed. Three theories have been advanced as to the part the silent corona discharge plays in the cracking process. It seems feasible to consider an electron moving through an appreciable distance in space and either carrying with it or dragging before it a mol. of H_2 in the vapor form. At last a surface is reached which permits no further movement of the oil vapor mol., but which readily absorbs the electron. The violent agitation of the vapors by the silent corona discharge is undoubtedly of prime importance in the complete cracking of the vapors, which facts tend to support the electronnech theory. The gasoline has high anti knock value and gives greater mileage than "doped" fuels. W. H. B.

Power cable technic. P. DUNNHEATH *Electrician* 106, 432-5(1931).—The best preventive of Pb-sheath corrosion is the provision of adequate, compounded, waterproof coverings over the Pb. The principal cause, if not the whole explanation, of ionization (rise of power factor voltage) in impregnated paper cables is found in the presence of voids, or spares not filled with solid or liquid dielectric. Sheath and armor effects and armoring of single core cables are discussed. For lower voltages the ordinary routine insulation resistance and pressure tests meet all requirements, but for higher voltages other tests are necessary. Breakdown tests seem more reliable than ionization stability tests, though each has its place. W. H. BOYNTON.

A new measurement of the temperature in the arc (V. ENGEL, STERNBECK) 3. Producing patterns in metals by photography and electrodeposition (Ger. pat. 515,900) 5. Purifying bitumen, etc., for use in "electrical contrivances" (Brit. pat. 339,470) 22. "Basing reagent" for muting lamp bulbs to bases (U. S. pat. 1,795,746) 18. Electric system for protecting boilers (U. S. pat. 1,796,715) 14. Extracting metals (Austrian pat. 121,523) 9. Allows for sheathing electric cables (Austrian pat. 121,523) 9. Float for determination of the specific gravity of accumulator electrolytes (Hung. pat. 101,225) 1. Treating metals in liquids as in an induction furnace (Brit. pat. 339,579) 9.

MÜLLER FRICH. A Laboratory Manual of Electro-Chemistry. Translated from 4th ed by H J T Fillingham. London. Geo Routledge & Sons, Ltd. 364 pp. 15s. Reviewed in *Chem. Trade J.* 88, 314 (1931)

WEJNARTH AXEL. Beiträge zur Kenntnis der elektrothermischen Zinkgewinnungsmethode. Stockholm. Wille Hefbergs Tryckeri Aktiebolag. 146 pp

Electric batteries. C DILLGER (tradig as Hamburger Batterie-Fabrik C. Dillger) and D SCHMIDT. Brit. 338,765, March 22, 1929. A battery of the Leclanche type has a depolarizer consisting of a double salt of NH_4 or an amine and Zn or other electropositive metal such as the double salt of ZnCl_2 and methylamine-HCl or *o* or *m* *p*-toluidine-HCl or Zn NH_4 chloride.

Electric batteries. LOUIS F J ROUSSEAU. Fr. 37,261, Feb. 21, 1929. Addn. to 648,569 (C. A. 24, 1584). Electrolytes of strong acid concn. are immobilized by fibrous masses of animal origin such as wool or felt or by dissolving an appropriate amt. of cellulose in the liquid.

Electric battery. MARIE IRLE NÉE GAUJACQ. Fr. 37,211, July 12, 1929. Addn. to 674,622 (C. A. 24, 2677). To prevent the Zn being covered with Cu or like metal resulting from the diffusion of the depolarizing liquid into the liquid which surrounds the Zn a metallic screen is placed in the layer of sepn. of the 2 liquids. A salt of Hg may be added to the depolarizing liquid, which salt by diffusion comes in contact and amalgamates with the Zn.

Dry batteries. KENO WOLF. Ger. 515,694, Oct. 25, 1929. The "dry" filling consists of colloidal H_2SiO_4 and concd. H_2SO_4 . A 15% soln. of H_2SiO_4 is preferred.

Dry-cell electric-battery assembly. ELMER J STROHL and EDWARD C. SMITH (to National Carbon Co.). U. S. 1,797,161, March 17. Structural features.

Storage battery. A G OWENS. Brit. 339,050, Oct. 9, 1929. Structural features.

Storage battery. JOSEPH L WOODBRIDGE. U. S. 1,796,818, March 17. Structural features.

Storage batteries. SOC. ANON. DES ACCUMULATEURS MONOPLAQUE. Brit. 338,921-2, Jan. 29 and Feb. 26, 1929. Structural features.

Storage-battery hydrometer (attached to the vent plug). CHARLES E LINTHARGER. U. S. 1,796,379, March 17. Structural features.

Battery container. EDWARD LYNDON. U. S. 1,796,034, March 10. Containers are made of permeable material non resistant to acids such as pressed wood pulp impregnated with prevulcanized rubber.

Depolarizers for galvanic batteries. NEUE ELEMENT-WERKE GEBR. HAAS & CO. Ger. 515,762, July 11, 1926. A depolarizer is made up by mixing together graphite and soot, compressing the mixt. and then mixing in MnO_2 .

Galvanic elements. FIRMA AUGUST ALEXANDER. Ger. 515,693, July 3, 1929. Details of arranging the C rods and metal caps in dry batteries.

Selenium cell. F H CONSTABLE and RADIOVOR PARENT, LTD. Brit. 339,245, Sept. 2, 1929.

Lead accumulator. JOHANN MATUSCHEK. Austrian 121,320, Sept. 15, 1930. The PbO_2 constituting the active mass of the pos. electrode is not exclusively generated during the manu. of the electrode, but is included in a substantial proportion in the mass from which the pos. electrode or both electrodes is or are to be formed. Other oxides of Pb may be present also, and the mass is preferably supported on unalloyed Pb. Details of the charging are described.

Protecting layer for the electrolytes of accumulators. GV DIAMANT Hung. 100,891, Feb. 20, 1929. A layer consisting of loose particles of solid materials is located on the surface of the electrolyte, e. g., charcoal particles impregnated with hydrocarbons or resins. This layer is afterward partly or totally solidified.

Dry rectifier with electrodes coated with metallic oxide. SIEMENS-SCHUCKERT-WERKE A-G. Ger. 515,529, Nov. 19, 1926. Plates of Cu are coated on one side with Cu_2O and on the other with Ag. Cf. C. A. 24, 27.

Electrodeposition of metals such as chromium, etc. HARRY SCHMIDT, GEORG BELWEIN and MARTIN HOSENFELD (to Metal & Thermit Corp.). U. S. 1,795,512, March 10. Solid adherent coatings are obtained by electrolytically depositing metals such as Cr at a temp. below their m. p. from a fused bath, the point of fusion of which is at least 100° lower than the m. p. of the metal to be deposited and which contains a B compd. such as mixed borates together with a compd. of the metal to be deposited.

Electrodeposition of chromium. WM S EATON. U. S. 1,795,481, March 10. The cleaned articles to be plated are immersed in an aq. electrolytic bath, the active ingredients of which consist solely of chromic acid and Na_2CO_3 in soln. in water.

Chromium plating LEON R. WESTBROOK (to Grasselli Chemical Co.). U. S. 1,795,459, March 10. A bath for Cr plating contains chromic acid in aq. soln. together with a small proportion of the SO_4 radical (which may be supplied by use of Na_2SO_4) and a small proportion of the neg. radical of molybdic or vanadic acid (suitably supplied by use of a molybdate or vanadate).

Electrodeposition of metals such as copper, nickel or cobalt on rotating cylindrical mandrels. S. O. COWPER-COLES. Brit. 338,457, May 15, 1929. Numerous details of app. and procedure are described for the deposition of the metal and its consolidation by the action of rollers which may be formed of glass or of metal such as ferro-Si. The mandrel is preferably of welded steel or iron sheet or turned and polished cast iron and to allow detachment of the deposit in tubular form the mandrel is first made passive by treatment with HNO_3 and then polished and treated with a stripping prep. An anode of Pb, C, ferro-Si or other material may be used and various details of electrolyte compns., etc., are given. Deposition of Zn and Ag also is mentioned.

Electrodeposition of metals on rotating mandrels. S. O. COWPER-COLES. Brit. 338,540, May 15, 1929. In a process such as described in Brit. 338,457 (preceding abs.), the surface of the rotating mandrel is prep'd to facilitate detachment of the deposit by forming on it a film of oxide or sulfide or by applying a film of Pb which is then polished with plumbago or other suitable polishing material. The surface of the mandrel is preferably formed of Cu or brass.

Electrodeposition of tin SIEMENS & HALSKER A. G. Brit. 338,488, May 16, 1929. Thick, coherent, firmly adhering deposits are obtained in the electrodeposition of Sn from alkali metal stannate or stannite solns., sol. anodes being used, by maintaining the free alkali content in the electrolyte (inclusive of that set free by hydrolysis, but not including alkali metal carbonate) below 15%. A Sn concn. of 0.05% or more may be employed, and if the Sn concn. is above 7%, the alkali content must be somewhat higher but should be not more than one-fifth of the Sn content. Cl. C. A. 25, 882.

"White gold" plating. DEUTSCHE GOLD- UND SILBER-SCHNEIDANSTALT VORMALS ROESSLER (Leo Nowack, inventor). Ger. 515,579, Oct. 15, 1926. Objects to be plated in "white gold" are formed of a ground metal which is given a thin layer of metal which then alloys with the ground metal to form the "white gold" surface. Thus, a degreased 18 c. gold object (Au-Ag-Cu) is nickled by electrolysis in the usual way. The Ni plated object is then immersed in a fused salt mixt. and heated to 700° for 20 min. This causes the Ni to alloy with the gold to form a "white gold" surface. Further examples are given.

Plating aluminum with other metals. M. K. DE TRAIRUP. Brit. 339,339, Sept. 3, 1929. Al surfaces are prep'd for plating with metals such as Ni, Cu, Ag or Au by a preliminary electrodeposition of an amalgam such as may be deposited from a bath formed of NaOH 200, sodamide 200, SnCl_4 40, ZnCl_2 10, $\text{Hg}(\text{NO}_3)_2$ 5, KOH 10 and water 20,000 parts.

Electroplating apparatus with a tumbling-barrel. CARL H. HANSEN. U. S. 1,795,336, March 10. Structural features.

Mold for making articles by electrodeposition. G. ROSENQVIST. Brit. 339,428, Oct. 24, 1929.

Liquid seal for gas chambers of electrical precipitation plants. LODGE-COTTRELL, LTD., and L. LODGE. Brit. 338,827, March 4, 1930.

Degasifying electrolytically deposited metals. JULIUS VON BOSSE. U. S. 1,795,484, March 10. Electrodeposited coatings of metals are freed from absorbed H. tending to cause embrittlement of the metal, by subjecting the metal coating to an elec. discharge in a vacuum tube in which the coating is made one of the electrodes.

Drying by electrolysis. N. ALTAT. Hung. 100,547, Sept. 3, 1929. Moisture is eliminated by electrolysis of material in its natural state or after grinding or pressing to briquets.

Electrolytic apparatus for removing salts from liquids. JEAN BILLITER. Austrian 121,039, Aug. 15, 1930.

Aluminum production by electrolysis of fused halide baths. I. G. FARRENINDE A.-G. Brit. 333,668, July 4, 1929. A fused mixt. which may be formed of NaCl 14, KCl 6 and AlCl_3 80% is electrolyzed at a temp. of 100° with a graphite anode and a rotating cylindrical cathode of graphite or metal (a compact deposit being obtained by use of a c. d. of 1 amp. per sq. dm. or even higher), or a mixt. comprising NaCl 18 and AlCl_3 82% may be electrolyzed at 160° with a W anode and an iron roller cathode. Various details of app. and procedure are described.

Electrolytic refining of aluminum. UKRAINSKII NAUCHNO ISSLEDOVATELNUII

INSTITUT METALLOV "UKRINMET" Russ. appl 30,500, July 24, 1928. A complex ammonium aluminum halide is used as electrolyte.

Magnesium. DOW CHEMICAL CO Brit. 339,504, Dec. 16, 1929. See U. S. 1,749,210 (C A 24, 2064)

Recovery of tin. SIEMENS & HALSKE A G Fr 37,425, Oct 7, 1929 Addn. to 650 579 (C A 23, 3173) In the electrolytic recovery of Sn from Sn residues not contg plumbite, bismutite or other substances acting in the same way, the electrolysis is carried out without the addn. of colloid after quantitative transformation of all the Sn into stannate form

Nitrogen fluoride OTTO RUFF Ger 518,202 Feb 2, 1930 NF_3 prepd by electrolyzing fused NH_4Hf , or by interaction of NH_3 and F_2 is freed from NH_4F and N_2O by passage over soda lime at 100° , followed by washing with warm acid KI soln. and fractionating

Coating metal articles with lead peroxide electrolytically. JOHANNES FISCHER (to Siemens & Halske, A-G) U S 1,795,941, March 10 Metal articles such as iron or steel are made anodes in an alk. bath contg Pb and at least one cyanide, sulfocyanide such as KCNS or dimethylglyoxime The bath is free from heavy metals other than Pb

Electrolytic condenser. JOSEPH SIFMAN (to Westinghouse Elec. & Mfg Co.) U S 1,795,900 March 10 Various elec. and structural features of condensers with film forming electrodes immersed in an electrolyte are described Cf C. A. 24, 1302

Introducing alkali or alkaline earth metals or oxygen into sealed glass vessels VEREINIGTE GLÜHLAMPEN UND ELECTRICITÄTS A G Austrian 121,033, Aug 15, 1930 This is effected by electrolyzing the wall of the vessel A part of the vessel is coated on the inside with a very thin metal layer, which constitutes one electrode, and the coated part is heated and supplied with current. The external electrode may be a fused bath, e. g., a salt of the metal to be introduced into the vessel Numerous details are given of the application of the method to the manuf. of cathodes for electric discharge tubes and photo-elec. cells Cf Austrian 120,121 (C A. 25, 1710).

Purifying inert gases in electric discharge tubes, etc. VEREINIGTE GLÜHLAMPEN UND ELECTRICITÄTS A-G Austrian 121,036, Aug 15, 1930 Alkali or alk. earth metals are introduced into the vessel contg. the gas by the electrolytic method described in Austrian 121,033 (preceding abst.)

Electric furnaces and electrolytic vats ALFRED FRANCHINI and PAUL GROSS Fr 695,216, Aug 27 1929 Method of mounting and moving the electrodes is described

Electric induction furnace. PORTER H BRACE (to Westinghouse Elec. & Mfg Co.) U S 1,795,827 March 10 Structural and elec. features

Electric induction furnace. NEVILLE R. DAVIS (to Westinghouse Elec. & Mfg Co.) U S 1,795,842, March 10 Structural features of a furnace, the lining of which has embedded in it a metal cage to effect its heating and drying by the heating from induction

Electric induction crucible furnace suitable for melting metals. PORTER H. BRACE (to Westinghouse Elec. & Mfg Co.) U S 1,795,926, March 10 Structural and elec. features

Coreless electric induction furnace suitable for melting metals. NEVILLE R DAVIS and CECIL R BURCH (to Westinghouse Elec. & Mfg Co.) U. S. 1,795,934, March 10 A metal conduit is arranged to cool an inductor coil which is adjacent and co-extensive in length U S 1,795,935 describes various details of procedure for sintering a desired portion of the lining of an elec. induction furnace.

Electric resistance furnace for heat treatments. SVEN P. AXELL (to Westinghouse Elec. & Mfg Co.) U S 1,795,819, March 10 Structural features of a muffle furnace provided with a conveyor

Electric resistance furnaces and ovens. F W. COFFE Brit. 338,816, Feb 11, 1930 Structural features

Electrical and mechanical control system for furnaces such as plural-hearth-tray rotary dumping furnaces. HARLAN S GANO (to Westinghouse Elec. & Mfg Co.) U S 1,795,945, March 10 Various structural and elec. features are described

Electrode holder for electric furnace SIEMENS-PLANAWERKE A-G. FÜR KOHLE FABRIKATE Ger 518,332, June 27, 1926

Electric regulating device for ovens for baking lacquered wire. ELECTRICAL RESEARCH PRODUCTS, INC Ger 518,078, Nov 16, 1927 The temp is regulated in accordance with the hardness of the coating on the wire leaving the oven.

Electrically heating the contents of high-pressure vessels. SIEMENS-SCHUCKERT-WERKE A-G (Günther Scharowsky, inventor). Ger 515,944, Aug 24, 1926.

Electric gas purifiers. METALLGESELLSCHAFT A-G (Walter Deutsch and Wilhelm Hoss, inventors). Ger 515,928, Nov 30, 1928. An arrangement for preventing damage to the electrodes by sparking is described. Cf. C. A. 25, 2061.

Apparatus for electrical precipitation of suspended particles from gases. INTERNATIONAL PRECIPITATION CO and LODGE-COTTRELL, LTD. Brit. 338,796, Jan 14, 1930. Structural features.

Apparatus for electrical precipitation of suspended particles from gases. LODGE-COTTRELL, LTD., and L. LODGE. Brit. 338,739, Nov 29, 1929. Structural features.

Electrical purification of gases such as carbon dioxide. SIEMENS-SCHUCKERT-WERKE A-G. Brit. 338,944, April 17, 1929. Various details are described for use of high-potential rectified a. c.

Use of silent electric discharges to facilitate catalytic gas reactions. FRITZ LECHLER (to International Precipitation Co.). U. S. 1,796,110, March 10. Apparatus and details of operation are described for promoting reactions by the combined action of contact substances and silent elec. discharge, in such manner that the elec. discharge acts not only to promote directly the reaction, but also to create an elec. wind in the gas so as to bring the gases into more effective contact with the contact substance. The process may be applied to the promotion of any reaction between materials in the gaseous phase, which is facilitated or promoted by the presence of a catalytic contact substance, for example to the oxidation of SO_2 to form SO_3 , as in the contact process of H_2SO_4 manuf. or to the chlorination or hydrogenation of hydrocarbons or other org. gases.

Spark-discharge apparatus for producing a bleaching gas for corn, etc. FRITZ STUFF. Ger 518,201, June 28, 1929.

Apparatus for heating "lyes and acids" by electrical resistances. VEREINIGTE ALUMINIUM WERKE A-G. Brit. 338,740, Nov 29, 1928. Tubular coils (suitably formed of Pb) serve as elec. resistances and communicate with the interior of the vessel, the contents of which are to be heated. These coils are placed around the vessel, and are in turn surrounded by heat insulating material.

Protection of electrical apparatus such as switches, transformers, choke, coils, motors or dynamos. MAX BUCHHOLZ. U. S. 1,796,746, March 17. A cell, the wall of which is of such a degree of porosity as will permit osmotic diffusion of the products of decompn. occurring in the insulation of the app. such as a transformer under abnormal conditions, is assocd. with a device such as an alarm device which is influenced by products thus passing through the wall of the cell. Various structural details are described.

Electric ozone generator. MARCUS P. OTTO. U. S. 1,796,214, March 10. Structural features.

Special loading of long submarine cables. V. E. LEGG (to Electrical Research Products, Inc.). Brit. 339,049, Nov 8, 1928. End sections of a cable are provided with loading of greater thickness than the central portion and which has a lower hysteresis loss, lower resistivity and a higher constancy of permeability for a wide range of magnetizing forces than the material of the central section loading. The central portion of the cable may be loaded with an alloy contg. Ni 45, Fe 23, Co 25, Mo 7, Si 0.47% or Ni 55, Fe 34 and Cr 11%. The end sections may be loaded with an alloy contg. Fe 27, Ni 45, Co 25, Mo 2.5 and Mn 0.35%.

Cut-out for series lamps. DANIEL S. GUSTIN and EWALD DIETZ (to the Canadian Westinghouse Co., Ltd.). Can. 309,241, Mar. 10, 1931. A cut-out medium for series lamps is produced by mixing powd. Cu with sodium silicate and dehydrating the mixt., then mixing the dehydrated mixt. with sodium silicate and applying pressure to produce a compact mass.

Electric incandescent lamp. LLOYD D. LOCKWOOD (to Westinghouse Lamp Co.). U. S. 1,795,785, March 10. Devices such as incandescent lamps are flushed with water gas during baking, in order to eliminate surface oxides on the metal parts. Cf. C. A. 24, 3958.

Decoration of electric incandescent lamps. K. MENESDORFER. Hung. 101,377, Nov. 12, 1929. Cellulose varnish is sprayed on the lamp as a base, and the decoration is produced by spraying various paints through stencils. The decoration thus produced does not crack off.

Coating for lamp bulbs. ALBERT F. LINDSTROM (to The Canadian Westinghouse Co., Ltd.). Can. 309,250, Mar. 10, 1931. A green coating compd. for the interior of lamp bulbs comprises sodium silicate, kaolin, hydrated chromic oxide and boric acid.

5—PHOTOGRAPHY

E. F. WIGHTMAN

Further advances in photographic science S O RAWLING *Science Progress* 25, 465-75(1931) —A review is given (with a bibliography of 23 references) of recent work on the chemistry of the photographic emulsion JOSEPH S. HIEBURN

Formation of photographic images on cathodes of alkali metal photoelectric cells. A R OLPIN AND G R STILLWELL *J Optical Soc Am* 21, 177-81(1931), cf C. A. 24, 5614 —Both neg and pos photographic images on the cathodes of K and Na photoelec. cells in a vacuum have been obtained by focusing an image of the object to be photographed on the cathode during the treatment with a suitable dielectric. The necessary optimum constitution of the surface structure will thus exist though perhaps only momentarily. The images are well defined and can be permanently "fixed" by proper treatment with various materials such as S vapor, air, O and H in the ratio of 9 to 1. HF and Br. The photoelec. sensitivity of the exposed areas decreases approx 30% during the formation of the image. There appears to be little difference in the sensitivity of the cathode area bearing the image and neighboring area after the image has been permanently fixed. The effect is undoubtedly photochem but the great variety of materials which can be used in treating the exposed surfaces complicates the chemistry of the phenomenon W. VANSELOW

Depression of density produced by the presence of bromide in the developer of photographic emulsions exposed to light and to x-rays A. P. H. TRIVELLI AND E. C. JENSEN *J. Franklin Inst* 211, 335-47(1931) —The density depressions of sol bromide in the developer were detd with 2 different emulsions for light and for x ray exposures with hydroquinone and with pyrogallol developer. Abegg's theory of the action of sol bromide in the developer was the best explanation of the results obtained. This theory was connected with the speck growth theory of the latent image, and the results obtained were explained on the basis of the size and the distribution of the specks. It was shown that there are 2 depression factors. (1) The difference in sizes of developable centers formed by exposure with sol bromide, gave a greater density depression for light than for x ray exposures. (2) The distribution of the developable centers throughout the thickness of the photographic layer will give, for light, a smaller density depression with sol bromide than for x ray exposures A. P. H. TRIVELLI

The question of the metallic ("Mir") silver content and germ formation in photographic emulsions H ARNOLD AND J EGGER *Z wiss Phot* 25, 178-90(1930) —1° O P emulsions, according to Valenta, increase their amt. of free Ag with age. For ordinary photographic layers the results are not const, which shows that the $Ag_2S_2O_8$ complex is not entirely separable from metallic Ag, so that Weigert and Lühr's data (C A 23, 4900) are too high and consequently also the previous data of the authors A. P. H. TRIVELLI

Cement for films (Hung. pat. 100,689) 23.

HIEBERT, L J. *A Manual of Photographic Technique* 2nd ed. London: Pitman. 133 pp. 2s 6d net.

WALL, E J. *The Dictionary of Photography*. 12th ed., revised by F. J. Mortimer. London: Hk. 681 pp. 7s 6d net.

Color photography. B E LUBOSHEZ *Brit* 339 511, Dec. 23, 1929. Various details are described of a color filter used with lenticular films.

Color photography. J N GOLDSMITH, T T BAKER, C BONAMICO AND SPICERS LTD *Brit* 339,238, Aug 1, 1929. In the production of regular pattern multicolor screens by means of dye repellent resists, the dyes in a non penetrative condition or medium are applied to intermediate surface layers only of a film support which may be formed of celluloid, cellulose acetate or other suitable cellulose ester or other compn.; e g., a powd. dye or finely divided dyed substance may be rubbed or spread on the film, or a dye such as cochineal in suspension or a dye in a viscous soln. such as castor oil or gum arabic soln. may be applied. Cf C A. 25, 1748.

Color photography. MULTICOLOR FILMS, INC. *Brit.* 339,323, June 4, 1929. Fixed Ag images are treated with a basic dye bath and then with a U toning and mordanting bath, in a process for producing color films having color component images on opposite sides which are simultaneously printed in register from color component nega

tives. The printed positive is developed, fixed and washed and that side bearing the images printed from the orange-red negative is toned blue by application to that side only of an Fe toning soln. which may be formed from ferric ammonium oxalate, $K_2FeC_2O_8$, NH_4Cl and HCl . The U toning bath may comprise K oxalate, U nitrate, $K_2FeC_2O_8$ and HCl . Various details and modifications of procedure are described.

Dye-impression color photography. J. E. THORNTON. Brit. 339,206, July 3, 1929. In producing 2-color pictures by dye-impression printing, a duplex dye-printing plate is prepd. as described in Brit. 339,321 having upon its composite faces a pair of dye images of complementary character, produced by exposing in a camera from one side only a sensitive film bearing complementary color sensitive emulsions on opposite sides with assocd. light filters, developing the images, treating them to render the gelatin dye-repellent in proportion to its Ag content, and dyeing with complementary colors the gelatin which remains dye receptive, also, pos. images are printed simultaneously from the duplex plate as described in Brit. 339,319, by pressure-contact upon 2 sheets of dye-absorbent non sensitive material provided with registration holds and placed one on each side of the duplex plate, and the printed dye positives are remented together in register and face to face. Numerous details of procedure are given both in this specification and in Brit. 339,319 and Brit. 339,321.

Screens for color photography. HAROLD WARD. Ger. 518,066, Jan. 25, 1929. See Brit. 322,432 (C. A. 24, 2681).

Screens for color photography. ANTON RASAVI and ALOIS GRAY. Austrian 120,858, Aug. 15, 1930. The screens are prepd. from fabrics woven from colored fibers, each fiber comprising a no. of differently colored filaments twisted together. Various methods of prepd. the screens are described.

Photographic films. I. G. FARBENIND. A. G. (Hermann Lummerzhelm, inventor). Ger. 515,776, Dec. 23, 1928. Addn. to 487,586 (C. A. 24, 1808). A roll film with opaque ends which are not affected by the developing bath, is coated with a layer of dye soln. in aq. or org. solvents which is opaque to actinic rays.

Photographic roll films. I. G. FARBENIND. A. G. Brit. 338,741, Nov. 29, 1928. The sensitive layer of a roll film with a protective paper backing is covered with transparent foil (which may be made of gelatin, regenerated cellulose, a cellulose ester, ether or ether-ester) colored to serve as a light filter when the film is exposed. Cf. C. A. 25, 887.

Roll films from cellulose acetate. KONAK-PATHÉ SOC. ANON. FRANÇAISE. Ger. 515,533, Apr. 20, 1927. The back of the film is colored by coating with an alc. soln. of magenta and aurine, with or without addn. of naphthalene green.

Photographic negatives. MAX WOLFRAM and EDWIN SCHAEFFER. Ger. 515,584, Dec. 20, 1929. Film negatives which can be easily detached from their glass, etc., carriers are prepd. by giving the transparent carrier a layer of rubber, gelatin or celluloid soln., before giving them the light sensitive layer.

Paste for over-exposed negatives. JOSEF PREIFFER. Austrian 121,139, Sept. 15, 1930. A paste for application to portrait and like negatives showing extremes of contrast comprises finely sifted kieselguhr 1 and tallow 3 parts, with a little $PhNO_2$ or wintergreen oil. For landscape and like negatives, only 1 part of tallow is taken, and coarse kieselguhr or tripoli is used instead of fine kieselguhr. The paste is rubbed on to the over-exposed parts, and acts mechanically by removing some of the excess of reduced Ag.

Photographic toned prints. RICHARD MICHEL. Austrian 121,141, Sept. 15, 1930. In the prepn. of toned prints on development papers contg. Ag salts by bleaching out, exposing and redeveloping, prints showing two or more tones are obtained by using a bleaching bath contg. an iodide and at least one other halide. A suitable bath is prepd. by adding 0.25-1.5 cc. of 1% KI soln. to a soln. of KBr 2, $CuSO_4$ 2 and citric acid 2 g. in water 100 cc. Addnl. details are given.

Photographic papers. CHARLES EMME. Fr. 37,318, Aug. 8, 1929. Addn. to 676,317 (C. A. 24, 2682). Blue varramine salt in aq. soln. is used as the diazo compd. to apply on the supports. A little sulfate of Ni or Al and a few g. of an org. acid may be added.

Apparatus for developing photographic paper by use of ammonia. KNUD MURCK (to Charles Bruning Co.). U. S. 1,795,344, March 10. Structural features.

Photographic layers. I. G. FARBENIND. A. G. (Walter Frankenburger and Georg Rössler, inventors). Ger. 518,064, Jan. 23, 1929. Light-sensitive substances and insensitive substances are suspended together in a colloidal material or a soln. thereof, and the suspension is distributed onto a support in the form of fine particles, e. g., by atomizing the suspension into a hot atm. so that a fine powder is obtained,

which is spread on the support. Different light-sensitive substances may be applied to a common support. Thus, for the prepn. of layers useful for color photography, three gelatin emulsions are prepd, each contg a silver halide sensitive to light in a primary color and a correspondingly colored pigment. The emulsions are atomized, and the resulting fine powders are mixed and applied to the support. Details are given also of the prepn of layers for making anaglyphs.

Anti-halation layers for photographic plates and films. I. G. FARBENIND, A.-G. (Hermann Durr and Wilhelm Schneider, inventors). Ger 518,363, Feb 18, 1930. These layers contain dyes obtained by condensing a substituted aromatic aldehyde with a quaternary heterocyclic base contg a reactive Me group. Such dyes are irreversibly decompd by the developing solns. A suitable dye is obtainable from quinaldine methyl bromide and *p*-dimethylaminobenzoaldehyde, or from quinaldine methyl sulfate and piperonal.

Printing on photographic layers. I. G. FARBENIND, A.-G. Fr. 37,317, Aug 12, 1929. Addn to 671,099 (C. A. 24, 2072). The 2 ends of the band of film are provided on the whole surface of the side of the layer with a color (obtained by the method of the prior patent) impermeable to actinic rays and insol in photographic treating liquid.

Photographic emulsions. V. L. MAKAROV. Russ appl 68,855, April 23, 1930. Photographic emulsion is sprayed on paper or plates with NH_3 or AcCl or with air satd with these compds.

Thioxanthonium or selenoxanthonium dyes for the preparation of photographic bleach-out layers. I. G. FARBENIND, A.-G. (Brunn Wendt and Alfred Fröhlich, inventors). Ger 518 065, Mar 1, 1930. Examples are given.

Photographic etching. ALEX B. DAVIS (to Keystone Watch Case Corp.). U. S. 1,797,210, March 17. A photographic resist is prepd by treating a hydrophobic colloid such as asphalt with a sensitizer comprising a soln of S chloride in CS_2 in a single step, removing the solvent, washing the treated colloid with an ether to increase its sensitivity to light, and dissolving in benzene. U. S. 1,797,211 describes forming a light-sensitive hydrophobic colloid by the reaction of a CS_2 soln. of an asphalt-base material with a CS_2 soln. of S chloride *en masse*, removing non reacted light-sensitive portions, placing a film of the light sensitive residue on an object to be etched, projecting a design on it, developing and etching. Cf C. A. 24, 2390.

Producing patterns in metal by photography and electrodeposition. OLIVE MILLER née ASSINDER and FREDERICK MILLER. Ger 515 900, Mar 23, 1930. A photographic plate having a sol light sensitive layer is exposed behind a diapositive of the pattern. An elec. conducting material in powder form is allowed to accumulate on the unexposed parts and this deposit is coated with an aq insol membrane, to which it adheres. The coatings are then removed and the conductor deposit of the desired pattern plated galvanically with metal.

6—INORGANIC CHEMISTRY

A. R. MIDDLETON

Coördinated compounds of the alkali metals. III. FREDERICK M. BREWER. *J. Chem. Soc.* 1931, 361-8, cf C. A. 20, 740.—The coördinated compds. of the alkali metals with salicylaldehyde (I) are prepd. The covalency nos. established for the metals in these compds are: Li, 4; Na, 4; K, 4 and 6 (the 6-covalent compd. is formed only with difficulty); Rb, 4 and 6 (only here the 6-covalent compd. is definitely more stable); Cs, 4 and 6 (only a slight difference in the stability of these 2 compds). Of the compds prepd those contg 4-covalent Na or K are by far the most stable. The compd contg 4-covalent Li ($\text{C}_7\text{H}_5\text{O}_2\text{Li}$, $\text{C}_{12}\text{H}_9\text{O}_4$, pale yellow solid) was obtained when dry LiOH was added to I and the mixt warmed gently. Both the 4-covalent compds of Na and K with I have been prepd by Hantzsch (*Ber.* 39, 3089 (1906)). Attempts to prep the compd contg 6-covalent Na were unsuccessful. The compd contg 6-covalent K ($\text{C}_7\text{H}_5\text{O}_2\text{K}$, $2\text{C}_7\text{H}_5\text{O}_4$, pale yellow leathery crystals) was obtained by adding concd. KOH soln. to I, dissolving the resulting solid in a mixt. of alc. and I, crystg., and washing with alc. A mixt. of the compds contg 4- and 6-covalent Rb was obtained as almost colorless crystals when powd. Rb_2CO_3 was dropped into warm L. On heating these crystals under reduced pressure a residue of rubidium salicylaldehyde ($\text{C}_7\text{H}_5\text{O}_2\text{Rb}$) was obtained. The compd contg 4-covalent Rb ($\text{C}_7\text{H}_5\text{O}_2\text{Rb}$, $\text{C}_{12}\text{H}_9\text{O}_4$, colorless crystals which melt indefinitely with decompn.) was prepd by dissolving Rb_2CO_3 in a min. of H_2O , adding I contg a small quantity of alc., warming until the CO_2 evolution ceases, cooling, recrystg. from alc. and washing with Et_2O . The compd

contg. 6-covalent Rb ($C_{11}H_5O_2Rb$, $2C_{11}H_5O_2$, colorless leathery crystals) was obtained (a) by allowing the product of recrystn in the preceding to stand in I for several hrs in the cold, and (b) by recrystg. from I contg a min of alc to effect soln. below 100° . The compd. contg. 6-covalent Cs ($C_{11}H_5O_2Cs$, $2C_{11}H_5O_2$) was prepd by carrying out the initial reaction without alc., recrystg. using sufficient H_2O to keep any excess Cs_2CO_3 in soln., filtering, washing with H_2O , keeping in I for several hrs., filtering and washing with alc. The compd. contg. 4-covalent Cs ($C_{11}H_5O_2Cs$, $C_{11}H_5O_2$, pale yellow cryst. solid which melts with decompn at about 165°) was prepd from the initial product by (a) washing with Et_2O until a slight decolorization begins, or (b) pptg. from an alc. soln. with Et_2O . In the compd. prepd, Li alone forms a true coordinated hydrate. *Rubidium benzoylacetone* ($C_{11}H_5O_2Rb$, $C_{11}H_5O_2$, bright yellow plates which melt to a red liquid at 131°) was prepd by dissolving Rb_2CO_3 in an excess of molten benzoylacetone, and recrystg. from hot toluene.

J. BALOZIAN

Reactions in the solid state at elevated temperatures. VIII. Reactions of columbium pentoxide and tantalum pentoxide with metallic oxides and alkaline earth carbonates. WILHELM JANDER AND HERMANN FREY. *Z anorg allgem Chem* 196, 321-34 (1931), cf *C A* 24, 4974.—This work involved a study of the reactions between BaO , SrO , CaO , MgO , ZnO , CuO , $BaCO_3$, $SrCO_3$, and $CaCO_3$ on the one hand and Ta_2O_5 and Cb_2O_5 upon the other hand. Heating curves, percent decompn and temp pressure curves were detd. by different methods. The following figures were given for the temp. at which a perceptible reaction commenced: $MgO + Ta_2O_5$ (485°), $ZnO + Ta_2O_5$ (500°), $MgO + Cb_2O_5$ (500°), $ZnO + Cb_2O_5$ (520°), $BaCO_3 + Ta_2O_5$ (450°), $SrCO_3 + Ta_2O_5$ (410°), $CaCO_3 + Ta_2O_5$ (400°), $BaCO_3 + Cb_2O_5$ (450°), $SrCO_3 + Cb_2O_5$ (420°) and $CaCO_3 + Cb_2O_5$ (380°). The following compds. were prepd for the first time: $2BaO \cdot Ta_2O_5$, $4BaO \cdot Ta_2O_5$, $2SrO \cdot Ta_2O_5$, $4SrO \cdot Ta_2O_5$, $4CaO \cdot Ta_2O_5$, $BaO \cdot Cb_2O_5$, $4BaO \cdot Cb_2O_5$, $SrO \cdot Cb_2O_5$, $4SrO \cdot Cb_2O_5$ and $5CaO \cdot Cb_2O_5$. $2CaO \cdot Ta_2O_5$ and $2CaO \cdot Cb_2O_5$ were also obtained.

H. STOERTZ

Obtaining anhydrous aluminum chloride from natural aluminum-containing raw materials. VICT. I. SPIRITVIN AND O. M. GLOSABEVA. *Z anorg allgem Chem* 196, 289-311 (1931).—The conditions for obtaining anhyd. $AlCl_3$ from kaolin earth and argillaceous earth-coal mixts were studied. Cl_2 and HCl gas were used as chlorinating agents. The raw material must first be freed of Fe by heating in a current of HCl gas at $400-900^\circ$, wood charcoal was the most successful form of C used. In the treatment with HCl gas an excess of C is essential, while with the treatment with Cl_2 this is not necessary, nor is gas velocity of great influence. In this reaction, silicic acid is likewise chlorinated. Thermodynamic considerations indicate that Al_2O_3 acts as a catalyst in promoting the chlorination of SiO_2 . The chlorination reaction begins at 600° and is accelerated at higher temps., the most favorable conditions occurring at about 800° . In an effort to diminish the chlorination of SiO_2 , HCl was added to the gas, but this likewise decreased the yield of $AlCl_3$. $SiCl_4$, however, when added to the Cl_2 , suppressed the chlorination of SiO_2 without appreciable effect upon the reactivity of the argillaceous earths.

H. STOERTZ

Crystallization of copper from molten cuprous chloride. V. SHIRONOV. *Z Elektrochem* 37, 80-2 (1931).—Cu crystd. out in dendritic form in the cooler portions of $CuCl$ unevenly heated to $370-470^\circ$. The reaction taking place in the melt was $2CuCl \rightleftharpoons Cu + CuCl_2$.

G. T. MOTOK

Some copper mercaptides and their reaction with carbon disulfide. W. E. DUNCAN, EMIL OTT AND E. EMMETT REID. *Ind Eng Chem* 23, 381-4 (1931).—By the interaction of solns of mercaptides in C_2H_5 with an aq. soln. of $(AcO)_2Cu$, the complete series of cuprous mercaptides from Me to n nonyl and from sec-Pr to sec nonyl was prepd. They react with S in a way similar to that observed for the Pb salts (*C A*, 24, 4758). They add CS_2 to form alkyl cuprous trithiocarbonates. No basic Cu mercaptides were formed.

LOUISE KELLEY

Hydrogen fluoride and its solutions. J. H. SIMONS. *Chem Reviews* 8, 213-35 (1931), cf *C A* 25, 861.—S. discusses the following: the prepn of anhyd. HF ; the technic of handling HF , the chem. and phys. properties of HF ; its mol. structure; the soly of substances in anhyd. liquid HF , which acts as a polar solvent very similar to H_2O ; the conductance of solns in anhyd. HF of salts, inorg. and org. acids, alcs, Me_2CO and Et_2O , the system C_2H_5-HF .

LOUISE KELLEY

The decomposition of Carborundum by a mixture of hydrofluoric and nitric acids. N. SLOMNECO. *Compt rend* 192, 361 (1931).—Finely divided SiC is decompd. by treatment with concd. HF and a little concd. HNO_3 for a period of 15 days. The liberation of gas marks the progress of the reaction.

F. W. LAIRD

Thiophosgene and some of its strange derivatives. ALBERT HUTIN. *Rev gén.*

mat. plastiques 7, 95-7(1931)—A few of the reactions of CSCl_2 are very briefly outlined

A. PAPINEAU-COUTURE

Preparation of hydrocyanic acid by oxidation of ammonium thiocyanate or thiocyanic acid with nitric acid. W. GLUD AND K. KELLER. *Ber. ges. Kohlen-techn.* 3, 395-419(1931). cf. *C. A.* 24, 2684.—In the oxidation of NH_4CNS with HNO_3 to get a good yield of HCN it is necessary to count on a 10% loss of NH_3 and NO , because of the fact that HNO_2 is formed which in turn forms NH_4NO_2 and this is decomposed into N and H_2O . If NH_4CNS however, is first converted into HCNS by treatment with H_2SO_4 , oxidation of the free acid to HCN can be brought about without this loss of N . The reaction proceeds according to the equation $\text{HCNS} + 2\text{HNO}_3 \rightarrow \text{HCN} + \text{H}_2\text{SO}_4 + 2\text{NO}$. A study made of the most favorable conditions of concn. showed that if a soln. of HCNS contg. about 20 g. per 100 cc. is treated with HNO_3 soln. contg. about 40 g. per 100 cc. a 95% yield of HCN can be obtained with very little loss of NO or HNO_3 . This represents 3-4 moles of HNO_3 per mole of HCNS , or a 100% excess of HNO_3 . Expts. were made on a larger scale with the following results. To manuf. 1000 kg. of HCN from 17% HCNS soln., 14,350 kg. of the HCNS soln., 13,950 kg. of HNO_3 (52.8%), 8.4 cu. m. of reaction water, 43.8 cu. m. of cooling H_2O and 10 kw. hrs. of elec. energy were required. This yielded, in addn. to the HCN , 2300 kg. of NO gas and 3820 kg. H_2SO_4 .

H. STROETZ

Constitution of cyanogen halides. II. Refractometric study of cyanogen chloride and iodide. ENRIGHE V. ZAPPÀ. *Bull. soc. chim.* [4], 47, 537-45(1930).—See *C. A.* 24, 2448.

E. M. SMOLES

Refractometric study of chromic acid reduction. GEORG GRASSER AND HIROSHI OHNO. *J. Faculty Agr. Hokkaido Imp. Univ.* 27, Pt. 2, 289-94(1930).—Changes in n accompanying reduction by a wide variety of reducing agents are described.

H. B. MERRILL

Liquid hydrogen sulfide as a reaction medium. JOHN A. WILKINSON. *Chem. Reviews* 8, 237-50(1931).—Liquid H_2S is discussed in regard to the following points: its prepn., its phys. properties, its action as a solvent for org. and inorg. substances, conductance of solns. of org. substances, particularly thiol and dithio acids, and inorg. substances in liquid H_2S , thiohydrolysis, or the metathetical reaction of H_2S with other substances such as the chlorides of P, As, Sb, Bi and esters of ClH_3COSiH , its reaction with various substances, such as liquid SO_2 , CaO , CaC_2 , nitriles and Grignard reagents.

LOUISE KELLER

The polysulfides of the alkali metals. II. Lithium. THOMAS G. PEARSON AND PERCY L. ROBINSON. *J. Chem. Soc.* 1931, 413-20. cf. *C. A.* 24, 4724.—An attempt to prep. Li polysulfides from soln. (by the action of S on Li_2S in alc., alc.- H_2O and H_2O solns. (I), and on LiHS in anhyd. alc. soln. (II)) and by dry methods (by the action of S on Li_2CO_3 (or LiOH) (III), and on molten Li (IV)). In I polysulfide solns. are obtained, which, however, could not be crystd. In II the primary product is Li_2S_4 , providing the S is in sufficient quantity, as demanded by $2\text{LiHS} + 3\text{S} = \text{Li}_2\text{S}_4 + \text{H}_2\text{S}$. On crystn. of the resulting soln. at room temp. $\text{Li}_2\text{S}_4 \cdot 2\text{C}_2\text{H}_5\text{OH}$ is obtained, while at the h. p. $\text{Li}_2\text{S}_4 \cdot \text{C}_2\text{H}_5\text{OH}$ is obtained. When S is used in smaller quantity than is demanded by the equation a mixt. of Li_2S_4 and LiHS results. In III a reaction first starts at 200° , giving thiosulfate and polysulfide, at 400° some S and thiosulfate are formed, and at 700° polysulfide, sulfate and small quantities of thiosulfate and sulfite are produced. In IV there is no reaction until 0.5 g. atom of S has been added to 1 g. atom of Li, when it takes place with explosive violence, giving principally Li_2S and small quantities of polysulfide. A study of the system $\text{Li}_2\text{S}-\text{S}$ shows that at atm. pressure only Li_2S (stable at its m. p., $900-975^\circ$) and Li_2S_2 (slightly decompd. at its m. p., 369.5°) are stable.

J. BALOZIAN

Autooxidation of phosphorus in carbon tetrachloride solution. BRUNO BLASER. *Ber.* 64B, 614-9(1931).—Besson (*Compt. rend.* 125, 1033(1897)) found that when dry air was conducted through suspensions of colorless P in CCl_4 there was formed a pale yellow ppt. from which he obtained P_2O . Blaser has studied this reaction quantitatively, using suspensions of finely divided P and an app. similar to that of Willstätter (*C. A.* 9, 308). With absorption of O , a pale yellow, voluminous ppt. began to sep. The velocity of oxidation was found to depend very much upon the temp. and the purity of the solns. Analysis of the product gave a P:O ratio of 2:3. The same product was obtained when P was used in excess. The compd. is not P_2O_4 , but is probably a new trioxide, $(\text{P}_2\text{O}_3)_x$. Its lack of color and the fact that no P can be extd. from it with solvents indicate that it does not contain impurities, yet its lack of the essential outward signs of homogeneity—m. p., b. p., crystal form and soly—demands further expts. to prove that it is a chem. individual. The new compd. was very hygroscopic.

Dried *in vacuo* it became white. It did not decomp. at 100° *in vacuo*; at a higher temp. it decompd. but did not melt. In air at room temp. the white compd. turned yellow and smelled of H_2S phosphide, at a higher temp. the yellow portions became red and P was formed. When some of the compd., suspended in CCl_4 , was added carefully to cold H_2O , a yellow ppt. was obtained which was apparently identical with the P_2O_5 described by Besson.

LOUISE KELLEY

The nature of hydrates of iron oxide. E. YA. ROBE. *J. Russ Phys.-Chem. Soc.* 62, 1443-52 (1930).—An investigation was made of temp.-dehydration curves of artificially prepd. iron oxide gel, of samples of granular lake-ore (Lake Pulo, Olonetzsk Gov.), xanthosiderite, brown ore (Ural, Bakal) and of gothite. The results agree with previous investigations in regard to the existence of 3 classes of natural hydrates of iron oxide, α -, β - and γ -.

S. L. MADORSKY

The mechanism of the formation of nitric esters. ROBERT C. FARMER. *J. Soc. Chem. Ind.* 50, 75-87 (1931).—F. discusses "pseudo acids," as designated by Hantzsch (*C. A.* 17, 2983), and shows that esterification is a property of the pseudo form and is complementary to the acidic or salt-forming function. Dil. HNO_3 solns. contain only small quantities of pseudonitric acid (I), HNO_3 solns. more concd. than 78% ($\text{HNO}_3 \cdot \text{H}_2\text{O}$) contain large amts. of I. Addn. of large amts. of H_2SO_4 may decrease the amt. of I, on account of the formation of nitronium sulfate, $(\text{HO})_2\text{N}(\text{SO}_3\text{H})_2$. F. shows that formation of cellulose nitrate occurs with HNO_3 of above 78% concn. in the absence of H_2SO_4 or with concd. HNO_3 in the presence of P_2O_5 or N_2O_5 . F. states that the H_2SO_4 used in nitrations does not have the "pioneering" action of first-forming sulfates, but acts only in dehydration. The formation of cellulose nitrate takes place best when about 10% of H_2O is present. With less than this, or in the presence of more than the usual amt. of H_2SO_4 , the nitrating action is suppressed, because of the formation of nitronium sulfate. The fact that with mixed acid principally nitrates are formed, the difficulty of formation of sulfur esters, and the use of alkyl sulfates in alkylation are explained by the reluctance of H_2SO_4 to form the pseudo acid. It is pointed out that the nitration of aromatic compds. is a function of the I present, and is suppressed except at higher temp. where nitronium sulfate is unstable, by the use of large proportions of H_2SO_4 . Numerous references are given to the literature on pseudo acids and nitration of cellulose.

G. R. YONGE

Contribution to the ebullioscopic study of complexes formed by the halides of cadmium and the corresponding alkaline halides. MAURICE O. HURN. *Compt. rend.* 192, 335-6 (1931), *cf. C. A.* 24, 5251.—Evidence is presented indicating the formation of $(\text{NH}_4)_2\text{CdCl}_4$, $(\text{NH}_4)_2\text{CdBr}_4$, and Na_2CdI_4 . For the chlorides and iodides the stability of these complexes decreases in passing from K to NH_4 to Na, while in the case of the bromides there is an inversion between Na and NH_4 . The affinity of these different complexes increases for the same alkali in going from Cl to Br to I. F. W. L.

Nitroso- and isonitrosoferropentacyanide derivatives from nitroprussiates. L. CAMBI WITH A. CAGNASSO AND T. RICCI. *Gazz. chim. ital.* 61, 3-13 (1931).—The present paper deals with the ferropentacyanides which are formed by condensation of alk. nitroprussiates with substances which contain $-\text{CH}_2-$ groups with mobile H. New reactions between nitroprussiates and aliphatic NO compds., cyanoacetic acid esters, PhCH_2CN and its derivs. and indole, which are described, confirm the mechanism of the reactions and the structure of the colored complexes already suggested by C. (*cf. C. A.* 7, 2551; 9, 451; 21, 1941), and data based on an x-ray examn. of aq. isonitrosoferropentacyanide derivs. of MeAc, acetylacetone, PhAc, AcCO_2H and indole confirm the structure of the complexes. EtNO (2.37 g.) added to solid K nitroprussiate (3.79 g.) in MeOH, cooled to 0° in an inert atm., EtOK (from 2.34 g. of K) added, the pptd. salt (which changes immediately from red to greenish gray) sepd. in an inert atm., dried, washed with MeOH and then with abs. Et_2O , dried *in vacuo*, dissolved in very dil. H_2SO_4 , extd. with Et_2O , evapd., and the residue recrystd. from CHCl_3 yields ethylnitrolic acid, m. 81° (decompn.). The same procedure was used with Na nitroprussiate (7.83 g.), $\text{NCCH}_2\text{CO}_2\text{Et}$ (3.39 g.) and NaOEt (from 1.38 g. of Na). The bright red salt when dried became yellow, and it also turned yellow in contact with water. The dry salt was decompd. as before, the product was extd. with Et_2O and purified from C_6H_6 , which yielded $\text{NCC}(\text{NOH})\text{CO}_2\text{Et}$ (*cf. Ber.* 42, 736 (1900)). With PhCH_2CN , the Na salt formed in abs. MeOH is red and is still more unstable than the preceding ones, changing to yellow. Decompn. as before yields an oxime which m. 105° and which after again solidifying m. 129° , probably as a result of isomerization. The compd. which m. 105° also m. 129° after recrystn. from water. It is therefore $\text{PhC}(\text{NOH})\text{CN}$ (*cf. Zimmermann, J. prakt. Chem.* [2], 66, 361 (1855)). The yield is 100%. Under the same conditions, $p\text{-O}_2\text{NC}_6\text{H}_4\text{CH}_2\text{CN}$ forms a dark red complex compd. which

changes rapidly to green. Decompd and dried, it forms an oxime, $C_{11}H_7O_2N_4$, m 95° , which confirms the data of Zimmermann (*loc cit*) for $p-O_2NC_6H_4C(NO)CN$. Similar results were obtained with other derivs of $PhCH_2CN$. The yields of isonitrosotriole in all cases were 85%, based on the nitroprussiate. Even in dil $PtOH$, indole condenses with nitroprussiates, with formation of intense blue complexes. K (0.85 g) in abs $MeOH$ (35 cc) added to K nitroprussiate (2.97 g) and indole (1.17 g) in abs $MeOH$ (50 cc) and the product washed with $MeOH$ and abs Et_2O yields the complex salt $K_3[C_{11}H_7ON_3Fe(CN)_5]$, probably contg the residue of nitroso- or isonitrosoindole. It is violet black and strongly hygroscopic and its concd aq solns, which are violet, become indigo blue when dilcd as a result of the formation of the acid salt $K_2[C_{11}H_7ON_3Fe(CN)_5]$, which is also deliquescent and is stable in water. Decompd did not lead to well defined stable complexes. Under the same conditions as in the prepn of $K_3[C_{11}H_7ON_3Fe(CN)_5]$, α -methylindole gives a reddish product which was not identified. β -Methylindole does not react, which makes it probable that in the formation of $K_3[C_{11}H_7ON_3Fe(CN)_5]$ and the reddish product above, the NO group enters in the β position of indole, as has been verified by Angel in the reaction of alc nitrites. The results of the x ray examn are shown graphically. The examn was limited to the visible field because of the tendency to decompose. The greater stability of $K_3[C_{11}H_7ON_3Fe(CN)_5]$ allowed measurements in the ultra violet region. With $MeAc$, acetylacetone and $AcCO_2H$, the salts prepd in water were compared with aq solns of the same salts prepd in abs $MeOH$. For comparison, the results with the $PhNO$ complex (cf Cambi and Ricci, *C A* 20, 1769, 24, 4231) are also given. The complex salts of $AcCO_2H$, $PhAc$ and indole when in dil aq soln, corresponding to the acid salts, show absorption bands similar to that of the pentacyanide from $PhNO$, viz, from $\nu = 490 \times 10^{-11}$ to 610×10^{-11} . Because of the displacement caused by various radicals united to the N of the NO group, this band is probably evidence of the similar chromophore group in the salts in question. The results show that alcoholates do not, as do alkalis, transform nitroprussiates into nitrocyanides, and that when the reaction can be carried out in water, the same products are obtained as in $MeOH$ (except for $(H_2N)_2CH$ and arylhydroxylamines, cf *C. C A* 24, 4231). The structure of the isonitrosoferropentacyanide derivs of nitroprussiates. A detailed discussion of the structural problem is included based both on the previous and present expts of C and on expts by others (cf *C loc cit*, C and Steg, *C A* 22, 2722, C and Ricci, *loc cit*, Iseff and Richarz, *C A* 22, 1291, Tarugi, *C A* 22, 1273, Zwickler, *C A* 24, 4239). In some cases the acid salts which are formed from primary nitroso complexes by weak acids or by great diln of their aq solns are relatively stable. They probably contain the oxime group thus $(NC)_5Fe-N \equiv C$, but in some salts at least the NO group

may be present $(NC)_5Fe \begin{array}{c} \text{OH} \\ | \\ NC \equiv \end{array}$ In this connection the facts are significant that

the complexes obtained from arylnitroso derivs of the type $[(NC)_5Fe \begin{array}{c} \text{O} \\ | \\ NPh \end{array}]Na_3$ are comparable in all their properties to nitroso- and isonitrosopentacyanides (cf *C A* 20, 1769, 24, 4231), and that the great stability of $K_3[C_{11}H_7ON_3Fe(CN)_5]$ is comparable only with that of derivs of arylnitroso compds, which it resembles insofar as it contains the NO united to the aromatic heterocyclic nucleus. The expts described in the present paper confirm the structures already suggested. In the reaction of nitroprussiates with $EtNO_2$, $NCCH_2CO_2Et$, $PhCH_2CN$, etc, it is probable that the salts initially formed, which are intensely colored, isomerize with loss of the characteristic chromophore group $(NC)_5FeN \equiv C$, which passes to $(NC)_5FeON \equiv C$, which in

turn undergoes cleavage, with loss of the oxime anion. With $PhCH_2CN$, the transformation would be $[(NC)_5FeN \equiv C(CN)Ph]Na_3$ (red) $\rightarrow [(NC)_5FeON \equiv C(CN)-$

$Ph]Na_3$ (yellow)

The color reactions of octacyanomolybdates. G A BARBIERI *Atti accad Lincei* 12, 148-53 (1930).—There are 2 series of compds. $K_2Mo(CN)_8 \cdot 2H_2O$ and $K_2Mo(CN)_8$, in which Mo has a valence of IV and V, respectively, which are analogous to ferro- and ferri-cyanides. To bring out this similarity further their reactions with Fe salts have been studied. $Mo(CN)_8^{2-}$ with ferric ions in acid soln gives a deep blue

C C. DAVIS

color, while $\text{Mo}(\text{CN})_6^{4-}$ does not. However, the blue is less stable, as no ppt. forms, the colored material being present in colloidal form. Addn. of Cu or Ag salts ppts the corresponding salt. Similarly, ferrous salts with $\text{Mo}(\text{CN})_6^{4-}$ give blue coloration (Turnbull blue). Uranyl acetate behaves in the same way, giving a red brown color with the IV-ion, but no color with the III ion. The Ag salt of IV is yellow, while that of III is red brown, corresponding to the ferro- and ferri-silver salts. A. W. C.

The decomposition tension of anhydrous uranyl nitrate and of anhydrous $\text{Th}(\text{NO}_3)_4$ (MISCIATELLI) 2.

CANALES TORO, MARÍA. Metales de la tierras raras. Santiago, Chile. Imprenta nacional. 48 pp.

A Text Book of Inorganic Chemistry. Edited by J. NEWTON FRIEND. Vol. VII, Part 2. Sulphur, Selenium and Tellurium. By R. H. VALLANCE, D. F. TWISS AND A. R. RUSSELL. Philadelphia. Lippincott. 391 pp. \$14. Reviewed in *Chem. Trade J.* 83, 314 (1931).

7—ANALYTICAL CHEMISTRY

W. T. HALL

An aid to calculations in gas analyses. JÖRGEN SCHMIDT. *Chem. Fabrik* 1931, 137-8.—A nomogram for calcg. H_2 and CH_4 is given. J. H. MOORE

Use of steam for Kjeldahl distillation of nitrogen. JESSE GREEN. *Ind. Eng. Chem., Anal. Ed.* 3, 160-1 (1931).—Constructional details and efficiency data are given. B. A. SOULR

Quantitative studies on the boric acid-alcohol flame test. WOLDEMAR STAHL. *Z. anal. Chem.* 83, 268-89 (1931).—See *C. A.* 25, 1455. W. T. H.

Notes on qualitative microchemical analysis. J. MIKA. *Bányász. Közl., Lapok* 63, 89-97 (1930).—Present qual. methods are described. S. S. DE FENÁLY

Importance of microanalysis. D. KISS. *Technika* 10, 146-7 (1929).—The development of microanalysis and a summary of its advantages over macroanalysis are given. S. S. DE FENÁLY

Microanalysis with an ordinary balance. I. Determination of nitrogen by micro-Dumas method. WM. J. SASCHKE. *Ind. Eng. Chem., Anal. Ed.* 3, 198-9 (1931).—A comparatively large sample of substance (0.1 g.) is ground up with a large amt. of finely pptd. CuO and an aliquot portion of the mixt. (about 0.2 g.) weighed on an ordinary analytical balance for analysis. The error in sampling in practice is less than 0.01% as calcd. by the Baule and Benedetti-Pichler formula. When the mixt. is weighed with an accuracy of 0.1 mg. on an ordinary balance, the substance under analysis has been weighed to 0.001 mg., which is the accuracy of the microbalance. N detns. were made of o-toluidine, acetanilide, azobenzene and benzamide by the Pregl-Dumas method with satisfying accuracy. L. T. FAIRHALL

Microacidimetric studies. I. J. MIKA. *Mikrochemie* 3, 143-64 (1931).—A theoretical discussion of the errors involved in microacidimetric work. The proper quantities of methyl red, bromothymol blue, phenol red, phenolphthalein, thymolphthalein and alizarin red for microchem. work are worked out. W. T. H.

Colorimetric investigations of indicators in presence of neutral salts. N. V. STOGWICK, W. J. WORBOYS AND L. A. WOODWARD. *Proc. Roy. Soc. (London)* 129A, 537-49 (1930).—The change of color exhibited by a definite concn. of an indicator in a soln. of given μ_H on the addn. of a neutral salt is investigated by a colorimetric method. An optical wedge is employed in conjunction with a Lundemann electrometer and a Rh photoelec. cell, the light absorption being measured by means of the wedge shift. The photoelec. effect enters into the method only as a null-point observation. The Tizard relation $K = [\text{H}] (C_1 - C)/(C - 1)$ is used, where K is the apparent dissocn. const. of the indicator acid, C the color of a slightly alk. soln. of methyl orange at a given concn. and C_1 the relative color, which is the ratio of the color of any other soln. to the unit color. A correction is made for the part of the H ions combined with the indicator anions to form undissoed indicator acid. Detns. of K are made in the absence of neutral salts with varying concns. of HCl , and a method is described for studying the possibilities that the neutral salt may alter the absorption bands of the colored forms of the indicator, and that it may affect the chem. equl. between these forms. The first possibility is studied independently of the second by dealing with solns. in which the indi-

color is completely in the yellow or the red forms. Measurements are also made of the dysoxon const. of methyl orange in the presence of varying concns. of several salts (NaCl, NaBr, NaO₂Cl, Na₂O₂, KCl and KBr), the concn. of the indicator being $N/40,000$. For NaCl, K rises to a max. between 0.05 N and 0.1 N and afterward falls. A series of measurements was also carried out with a const. quantity of AcOH-NH₄OH buffer, instead of HCl, at an indicator concn. of $N/53,000$, and the results were of the same type for all salts: the value of $[H^+_{free}]$ rises with increase of salt concn., passing through a max. at about 0.5 N . The theoretical significance of these results is discussed.

A. L. HENVE

Determinations by Wood light. A. G. NASINI AND P. DE CORI. *Atti III congresso naz. chim. pura applicata* 1930, 668-78.—The use of eosin, coumarin, umbelliferone, quinine, naphtholsulfonic acids, fluorescein, acridine, dichlorofluorescein and β -methylumbelliferone as fluorescent indicators for dark solns. is reviewed. The pH of the soln. affects the intensity of the light emitted. Detns. were made on black liquor, sulfuretted oils, wines, beers, acid FeSO₄ solns. and com. indigo carmine, certain advantages as well as difficulties in the use of fluorescent indicators being shown.

E. M. SYMMES

Volumetric determination of selenium. B. ORMONY, *Z. anal. Chem.* 83, 338-41 (1931).—Benesch (C. A. 23, 790) proposed a method for detg. Se based upon titrating a Na₂S soln. with KCN. It is pointed out that Se forms the compd. Na₂SeSe₃·5H₂O when it dissolves in Na₂S, instead of a colloidal soln. of Se as B. assumed. It seems probable that difficulties will be encountered in practice in attempting to det. Se by this method.

W. T. H.

The microchemical determination of strontium. L. DE ZOMBOY, *Technika* 10, 147-9 (1929).—The Sr (4-5 mg.) is pptd. as SrCO₃ and weighed after drying at 150°; or the ppt. can be treated with H₂SO₄ and weighed as SrSO₄.

S. S. DE FINALLY

Separation of iron-titanium-aluminum in tartrate solution. E. SCHWAB VON BECKHOF, *Z. anal. Chem.* 83, 345-60 (1931).—A satisfactory sepn. of Fe, Al and Ti (corresponding to 0.084 g. Fe₂O₃, 0.018 g. TiO₂ and 0.07 g. Al₂O₃) was accomplished by first reducing the Fe⁺⁺⁺ with H₂S in acid soln., making slightly ammoniacal and pptg. FeS with H₂S, making acid with H₂SO₄ and pptg. Ti with cupferron, and finally pptg. Al with oxalic acetate in ammoniacal soln. The presence of V caused trouble.

W. T. H.

Determination and separation of lead and bismuth by the volumetric filtration method. HANS TH. BUCHNER AND P. W. MEIER, *Z. anal. Chem.* 83, 352-61 (1931).—B. and M. (C. A. 25, 47) have shown the advantages of carrying out quant. pptns. with measurement of the vol. of reagent required to effect complete pptn. as detd. by filtering the soln. and testing with more reagent and also with a soln. contg. the ion to be pptd. For detg. Pb, pptn. as PbSO₄, as Pb₃(PO₄)₂, as PbCrO₄ and as PbSeO₄ was tested, and the last two methods proved suitable for this kind of an analysis. For pptn. of Bi, a 0.015-0.02 M soln. of H₂SeO₄ is suitable. The Bi selenate is formed in 0.05-0.08 N HNO₃ soln., but HNO₃ prevents the pptn. of the corresponding Pb salt. For the detn. of Bi and Pb, therefore, make the soln. 0.05-0.08 N in HNO₃ (0.15 g. Bi or less) and the vol. of soln. about 100 cc. Heat to 70° and add the H₂SeO₄ soln. with shaking. Toward the end of the reaction, heat to boiling. After the detn. of the Bi, add an excess of NaOAc, heat to 80-90° and ppt. the Pb (0.1 g. or less) with the same H₂SeO₄ soln. or with K₂Cr₂O₇ soln. The results obtained in 8 analyses all agreed within 0.23 mg. (1.5% of the Pb content).

W. T. H.

Separation of magnesium from potassium and sodium in the analysis for cations. G. KALKINSKI, *J. Russ. Phys.-Chem. Soc.* 62, 1355-6 (1930).—After the pptn. and filtration of MgNH₄PO₄, the excess PO₄⁻⁻⁻ is removed by adding CaCO₃, and the filtrate from the Ca₃(PO₄)₂ ppt. is divided into 2 portions and sep. tests made for Na⁺ and K⁺.

S. L. MADORSKY

Microdetermination of calcium and magnesium in the presence of each other. K. L. MALYAROV, *Mikrochemie* 3, 132-5 (1931).—If not over 10 mg. of CaO or MgO (as carbonate, oxide or hydride) and only very little, if any, alkali carbonate is present, a satisfactory sepn. and detn. can be accomplished by taking advantage of the fact that CaO is quite sol. in water and MgO is not appreciably sol. Ignite the sample for some time, preferably over the blast lamp, cool in a desiccator over soda lime or KOH and then dissolve the CaO by means of 60-80 cc. of hot water free from CO₂. Filter and titrate the filtrate with 0.05-0.02 N HCl with methyl orange as indicator. Dissolve the residual MgO in a measured vol. of the HCl and titrate the excess acid with NaOH.

W. T. H.

Determination of calcium and magnesium by titrating in the same solution. K. L.

MALYAROV *J. Russ Phys. Chem. Soc.* 62, 1529-31 (1930).—Ca is pptd. by means of oxalic acid, and then, without filtration, Mg is pptd. as hydroxide from the same soln. by means of KOH. The two can then be detd. microanalytically, without sepn., in the following way: Mg is detd. by titrating with H_2SO_4 , and Ca is detd. by titrating with $KMnO_4$. The soln., however, must be free from NH_4 salts, and the amt. of MgO must not exceed 0.01 g.

S. L. MADORSKY

Detection of gold, palladium and silver with dimethylaminobenzylidenerhodanine. F. FRIGL, P. KRUMHOLTZ AND E. RAJMAN. *Mikrochemie* 3, 165-73 (1931)—p-Di-

methylaminobenzylidenerhodanine, $S \cdot CS_2 \cdot NH_2CO \cdot C(CH_3)_2 \cdot NMe_2$, can be used to advantage for the detection of Au, Pd and Ag. To detect Ag in the presence of Au, Pt, Pd and Hg, take a drop of the soln. on a spot plate, mix it with a drop of 10% KCN soln., add 1 drop of the reagent dissolved in alc. and make acid with a few drops of HNO_3 ; a violet ring forms if Ag is present. Cu must be absent, as it also gives a color reaction with the reagent. 0.001 mg of Ag can be detected in the presence of 1000 times as much Hg, 4000 times as much Au and 300 times as much Pd. Au, like Ag, gives a colored ppt. with the reagent. The reddish violet ppt. can be used as a test in a test tube or on the spot plate. After adding the reagent, its excess can be removed by shaking with ether, and a red skin will form at the layer between the 2 solns. Palladous salts give a violet ppt. Of the other Pt metals, Pt and Ir also give tests with the reagent, but not so readily; so it is possible to test for Pd in the presence of other Pt metals. Pd can also be detected in the presence of Ag if the latter is first converted into $[AgBr_2]^-$. If Au is present, it can be reduced to metal by treatment with alkali nitrite in a soln. kept neutral by the addn. of an excess of $CaCO_3$.

W. T. H.

Determining molybdenum in steel. H. C. WERRICK AND C. H. MCCOLLAM. *Heat Treating and Forging* 16, 1145-6, 1155 (1930)—In the detn. of Mo by the $PbMoO_4$ method 2 sources of error may appear: (1) the interference of elements other than Mo, (2) errors found in the method itself. If Al or V is indicated by sep. analysis, either the sulfide pptn. or the colorimetric method (Maag and McC., *C. A.* 19, 1675) should be used. In carrying out the procedure frequent blanks must be made on the reagents to check their purity; the Pb and the NH_4 acetate solns. should be reasonably fresh; the NaOH should be of high purity, and the correct amt. of HCl should be added.

J. BALOZIAN

Detection of fluorine in plants and soils. PAUL RECKENDORFER. *Mikrochemie* 3, 126-31 (1931)—The method depends upon the combustion in a calorimetric bomb with O_2 at a pressure of 25 atm. About 2 g of the sample was compressed to a small pill, and this pill was covered with a cotton wick. A little 20% KOH soln. was also placed in the bomb. The combustion was effected as in the detn. of the heating value of coal. After the combustion the ash and the KOH soln. were examd. separately for F, the test depending upon volatilization as SiF_4 by treatment with SiO_2 and H_2SO_4 , and carrying out the molybdate-benzidine test for the sol. Si compd. By means of soil and plant tests it was shown in one case that the F content of a plant was due to fumes and not to absorption from the soil.

W. T. H.

Rapid determination of small quantities of boric acid by the intensity of the flame coloration. WOLFGANG STAHL. *Z. anal. Chem.* 83, 340-4 (1931); *Acta Univ. Latviensis Kim. Fakultat Series* 1, 401-7 (in Lettish), 407 (in German)—With the aid of turmeric paper and standard solns., Bertrand and Agulhon (*C. A.* 8, 2005) detd. 0.0005 to 0.1 mg of B. By means of the green flame test with $Mg_2B_2O_7$ it is easy to det. 0.3 mg. of B_2O_3 or more with an accuracy of 0-15%. For this purpose it is necessary to have a colorimetric scale, produced with standard solns. of known B content, 3 Bunsen burners of the same size, and an app. suitable for introducing equal vols. of air through the soln. being tested and through 2 standard solns. The best mixt. consists of 100 vols. of MeOH and 20 vols. of H_2SO_4 . Directions are given in detail and the results of many tests are included.

W. T. H.

Titration of sulfate in the presence of ferric ions. L. DE ZOMBORY. *Technika* 10, 192-3 (1929)—If 1-3 cc. of dil. H_3PO_4 is added to a soln. contg. SO_4^{--} and Fe^{+++} , then an excess of 0.5 N $BaCl_2$ soln. should be added. The excess $BaCl_2$ can be detd. by titration with 0.2 N alkali or NH_4 sulfate soln. in the presence of 5-10 drops of 0.25% Na rhodizonate soln. as an indicator. The red color changes to a greenish white at the end point.

S. S. DE FINÁLY

Method for the estimation of iron in biological material. ROBERT HILL. *Proc. Roy. Soc. (London)* B107, 205-14 (1930)—Addn. of $\alpha\alpha'$ -dipyridyl to a soln. of a ferrous salt between pH 3.5 and 8.5 yields an intense red complex ferrous ion. The color is not influenced by the presence of other metals unless they are present in great excess over the Fe

Their presence may be overcome by use of sufficient excess of the reagent to insure complete transformation of the Fe into its dipyrindyl complex Fe^{+++} , except in high concn, gives no color and does not interfere with the reaction for Fe^{++} . Both Fe^{++} and Fe^{+++} may be detd by addn of dipyrindyl and measurement of the intensity of the pink color both prior to and after reduction. The best reducing agent for this purpose is $\text{Na}_2\text{S}_2\text{O}_4$ prepd as follows. To an almost satd soln of $\text{Na}_2\text{S}_2\text{O}_4$ at 40° add sufficient $\alpha\alpha'$ -dipyrindyl to produce the max red color (traces of Fe^{++}), filter rapidly on a Buchner funnel, and ppt $\text{Na}_2\text{S}_2\text{O}_4$ from the filtrate with alc. Let stand 20 min, collect the ppt. on a filter wash with 70% alc until colorless, then with 97% alc. Boil with 97% alc. for 10 min, filter immediately and transfer the salt at once to a vacuum desiccator contg H_2SO_4 . Hydrazine hydrate may also be used as a reducing agent at a temp of 40° . $\alpha\alpha'$ -Dipyrindyl does not react with reduced hematin nr pyridine hemochromogen and does not remove Fe from hematin, it inhibits the catalytic reactions of Fe ions. Iodides, tungstates, alkaloidal reagents and $\text{Fe}(\text{CNS})_3 + \text{HCl}$ tend to interfere with the development of the red color usually by causing formation of a ppt. Permanent standards are prepd by scaling 5 cc of a standard soln in a tube 1.3 cm in diam and 8 cm long. The standard solns are prepd from 2 stock solns, 0.01 M ferrous ammonium sulfate, and 0.03 M $\alpha\alpha'$ -dipyrindyl HCl , these solns are equiv from the viewpoint of the colored complex ion. The most concd standard is 0.0001 M ferrous dipyrindyl, contg twice the theoretical amt of the dipyrindyl group, it is prepd from the 2 stock solns, acetate buffer and a trace of $\text{Na}_2\text{S}_2\text{O}_4$. From this standard are prepd 23 addn standards, contg from 0.0051 to 0.00056 mg Fe^{++} per cc, acetate buffer soln being used as the diluent. SO_2 is passed into the Standards prior to scaling in order to insure their permanence. The p_n of unknown solns is brought within the proper limits by addn of a soln of Fe free AcONa contg AcOH . The ferrous dipyrindyl may be adsorbed by proteins in neutral, slightly alk. or even slightly acid solns, this usually may be prevented by use of AcONa plus SO_2 , or of 30% alc. These procedures may be used for the detn of Fe in biol material without previous incineration. Baker's yeast contained 0.0016% non hematin Fe^{+++} , hen egg yolk 0.0085% Fe^{+++} probably as colloidal $\text{Fe}(\text{OH})_3$. At p_n 7.3 low concns of $\alpha\alpha'$ -dipyrindyl inhibit the catalysis of the oxidation of cysteine by Fe but have no influence on the action of Cu on this reaction.

JOSEPH S. HEPBURN

Determination of gold in animal substances. S. TUKATS AND M. LEINZINGER. *Magyar Gyógyszerészeti Társaság Értéktöze* 6, 43-54 (1930).—Treat the sample with concd HNO_3 evap and heat the residue in an elec oven to about 300° . To the ash add 3 cc HCl and 30 cc of CaOCl_2 soln and evap to dryness. Add 1-2 drops of 25% HNO_3 , 2 cc of fresh Cl_2 -water and 5 cc of water. Add 1-2 drops H_3PO_4 and boil 10-15 min to remove Cl. The soln is now ready for an electrolytic, gravimetric or iodometric detn.

S. S. DE FINÁLY

The Nessler reagent and its action on reducing sugars. LIVIO PAVOLINI. *Chim ind agr biol* 7, 39-40 (1931).—Instead of Fehling soln, which is blue and does not permit a good observation of the color changes, P. proposes Nessler soln. (alk. soln of K_2HgI_4) for the detection of sugars. 5 cc. of cold soln contg. 0.1-0.2 g sugar are treated with 0.5 cc of reagent and shaken. a yellow or green yellow ppt is obtained with glucose, lactose, fructose and dextrin. Sucrose does not give the test. G. A. B.

Microdetermination of glucosides in plant materials, with emphasis on the difficulties. ANNELESE NIEMHAMMER. *Mikrochemie* 3, 130-42 (1931).—Rosenthaler (C. A. 24, 3531) has indicated the advantages of sublimation tests for the identification of glucosides in plant materials. His work is confirmed by studies on the bark of the wild horse chestnut contg. esculin, seeds of the yellow wattle and of eow wheat contg. rhamanthin, bark and leaves of the elder contg. syringin, leaves of the common leek contg. saponarin, digitalis plants contg. digitonin and bark of the *Salix caprea* contg. salicin. The glucosides to be sure, undergo slight decompn during the sublimation, but the sublimate usually yields crystals which can be identified, particularly after treatment with Br KBr soln.

W. T. H.

The determination of arsenic in arsenobenzene. UGO CAZZANI. *Giorn farm chim* 79, 62-4, 67-70, 73-4, 77, 107 14, 117-8, 163-6, 169-72, 175-8, 181-3, 186-7 (1930), cf C. A. 20, 263.—A complete review is given of the various methods employed in the detn. of As in arsenobenzene. The methods are described in detail and their merits and defects are discussed. Lehmann's method (with slight modifications), De Myttenaere's method, and the bromometric method of Kircher and Ruppert are considered to be the best methods, analytical evidence is given in support of this view. The following modification of Lehmann's method is suggested. Mix 0.2 g of arsenobenzene with 1 g of finely powd KMnO_4 , add 5 cc of 30% H_2SO_4 slowly, with gentle

agitation. Allow to stand 10 min., shaking occasionally. Then add 10 cc. of concd. H_2SO_4 and some H_2O_2 soln. Add 5 cc. of H_2O_2 in excess after the liquid has become limp and colorless. Dil. with 25 cc. of H_2O and gently boil the liquid until white heavy fumes appear. After cooling add 50 cc. H_2O , cool again, add 2.5 g. of powd. KI , cover the flask with a watch glass and shake until the KI is dissolved. Allow to stand in the dark for 1 hr. and then titrate with 0.1 N $\text{Na}_2\text{S}_2\text{O}_3$ without starch. Run a blank in the same way. C prefers the following procedure. To 0.20 g. of substance add 10 cc. of H_2O_2 soln. and 10 cc. of concd. H_2SO_4 . Heat gently on the sand bath until more gas is evolved. Remove from the sand bath, add 5 cc. more of H_2O_2 soln. and heat gently until no more gas is evolved, then heat strongly for 10–15 min. Dil. the resulting colorless liquid with 50 cc. H_2O , add dropwise 1% KMnO_4 until a rose color persists. Decolorize with oxalic acid soln., cool, add 2.50 g. of powd. KI and shake the soln. in a covered flask until the KI is dissolved. After 30 min. in the dark titrate with 0.1 N $\text{Na}_2\text{S}_2\text{O}_3$ soln. without addn. of starch. Then add 75 cc. of said Na_2CO_3 soln. and about 10 g. of NaHCO_3 and titrate with 0.1 N I until a light yellow color persists. Run a blank test in the same manner. Comparative results are shown with various procedure and 2 pages of references are given. G. SCHRÖCH

Critical study of the method proposed by G. Florence for finding alkaloids in viscera. N. J. JOANID *Bull. soc. chim. bel.* 12, 1001–13 (1930), cf. *C. A.* 21, 3905.—Recovery of strychnine, morphine and cocaine from fresh and putrefied beef was much lower when the method proposed by F. was used, instead of the method of Kohn-Abrest (modification of Stas-Otto method). Recovery from the organs of a dog poisoned with strychnine was only 0.25 that recovered by the older method. C. G. KING

Detection of semen in legal cases. J. PELTZER *Chem.-Ztg.* 55, No. 7, 70 (1931).—Microscopic examn. of an aq. or H_2O_2 suspension may disclose characteristic elements or spermatozoa. The suspected spots or flecks are moistened with H_2O_2 . If semen is present, much foaming occurs. Microscopic mounts may be stained with 2% aq. soln. of eosin. If the mount is pos., long, lance-like characteristic crystals (of Florence) appear. The addn. of 1- KI soln. to the mount in very small amt. colors the crystals chocolate-brown at first, gradually the crystals disappear but are recrystd. by the addn. of a little more of the I soln. The presence of spermatozoa or of the Florence crystals is pos. evidence of the presence of semen. C. R. FELLERS

Titration of ethylenic nitriles. G. HEIM *Bull. soc. chim. Belg.* 39, 458–61 (1930).—Ethylenic nitriles, with a double bond in the α - β position, add I_2 much more slowly than nitriles with a β - γ double bond. If the following conditions are respected, a mixt. of both forms can be titrated by bromination, because for practical purposes the α - β forms are not affected. Dissolve the nitriles in enough CHCl_3 to obtain 1/40 mol. in 100 cc. of soln. The brominating soln. is a 0.1 N Br aq. soln. contg. 100 g. KBr/I . For titration, add twice the quantity of Br soln. needed for the amt. of nitrile, shake vigorously for 10 seconds to emulsify, add a 10% KI soln. and titrate the I liberated. A. L. HENNE

Three new reactions of alanine. JUAN A. SANCHEZ *Semana med.* (Buenos Aires) 1931, I, 651–3.—(1) Heat 0.01 g. alanine with 20 drops of 1% KMnO_4 soln. to 100° for 1 min., cool, add 0.05 g. oxalic acid and then to the colorless liquid 2 cc. EtOH , 0.02 to 0.03 g. o-nitrobenzaldehyde and 10 drops of 30% NaOH . Shake with CHCl_3 , the CHCl_3 takes a blue color. The reaction is explained as being due to the formation of indigo. (2) Heat alanine with NaOCl for some seconds, add 2 drops of 30% NaOH and by drops a soln. of 1 in KI . A ppt. of CHI_3 is formed. The intermediate formation of AcI is supposed. (3) Heat dry alanine in a closed tube and dissolve the condensed vapors in dil. HCl . The soln., which contains EtNH_2 , gives with the reagent of Boucardat a reddish brown ppt., with Dragendorff's reagent a red one and with S's molybdic reagent a white ppt. A. E. MEYER

Determination of benzene in alcoholic solutions. GY. GRÖN AND E. FALTING. *Magyar Chem. Folyóirat* 36, 166–9 (1930).—Small quantities of benzene may be detd. by means of the characteristic absorption in the ultra violet spectrum. Extinction coeffs. of alc. benzene solns. were detd. by a Hilger's spectrograph in layers of various thickness, and it was found that the law of Beer- Lambert holds. Sensitivity was found to be 0.01 g. per l. S. S. DE FINÁLY

Color tests for some saturated and unsaturated carboxylic acids. L. ECKERT. *Magyar Gyógyszerészeti Társaság Értesítője* 7, 121–4 (1931).—Color tests with different aldehydes were tried. Saturated monocarboxylic acids with few C atoms do not give tests with aldehydes in alc. H_2SO_4 solns. Acids with more C atoms, e. g., palmitic and stearic acids and unsatd. acids, e. g., oleic and ricinoleic acids, show, on the contrary, striking color reactions. S. S. DE FINÁLY

Examination of a new reaction of acetic acid. J. FERDÖS. *Magyar Gyógyszerészeti Társaság Értesítője* 6, 372-6 (1930).—The color test with La nitrate described by Krüger and Tschirch (*C. A.* 23, 5129) was studied. The minimum quantity of AcOH giving a blue color was found to be 0.1 mg. as stated by K. and T. The test fails if over 1 1/2 times as much lactic acid is present, no test was given in 1% soln. by several mono-, tri- and amido-substituted acetic acids, acetonitrile, aniline acetate, Et acetate, or iso-butyl acetate. NaOAc gave the blue color. The following conclusions are drawn: (1) Substitution in the Me radical of acetic acid, (2) esterification of carboxyl by organic radicals and (3) substitution of a carboxyl group prevent the reaction, but (4) inorganic salts of AcOH give a positive test. S. S. DE FINÁLY.

Allen's Commercial Organic Analysis. 5th ed., revised. Vol. VIII. Glucosides, Non Glucosidal Bitter Principles, Lixzymes, Putrefaction Bases, Animal Bases, Animal Acids, Cyanogen and the Cyanogen Halides, Proteins and the Digestion Products of the Proteins. Edited by S. S. Sadtler, E. C. Lathrop and C. A. Mitchell. Philadelphia: P. Blakiston's Son & Co. 733 pp. \$7.50. Cf. *C. A.* 22, 1303.

BAILEY, E. H. S., and CARY, HAMILTON. Qualitative Analysis. 8th Ed., revised by Paul V. Faragher. Philadelphia: P. Blakiston's Son & Co. \$2.

BRISCOE, HERMAN T. Qualitative Chemical Analysis. Principles and Methods. New York: D. Van Nostrand Co. \$2.25.

CLOWES, FRANK AND COLEMAN, J. BERNARD. Quantitative Analysis. 12th ed. Philadelphia: P. Blakiston's Son & Co. \$5.75.

LUNGE, G., and KEANE, CHAS. A. Technical Methods of Chemical Analysis. 2nd ed., revised and edited by C. A. Keane and P. C. L. Thorne. Vol. III. London: Gurney & Jackson. 698 pp. £3.31, net. Cf. *C. A.* 23, 580.

SUTTON, FRANCIS. Volumetric Analysis. 11th ed., revised by W. Lincoln Sutton and Alfred E. Johnson. Philadelphia: P. Blakiston's Son & Co., Inc. \$9.

8—MINERALOGICAL AND GEOLOGICAL CHEMISTRY

EDGAR T. WHERRY AND J. P. SCHIAIREX

Röntgenographic examination of remelted algonite and whitneyite. Supplement. FELIX MACHATSCHKE. *Centr. Mineral. Geol.* 1929A, 371-3, *Chem. Zentr.* 1930, I 957, cf. *C. A.* 25, 1450.—The expts. of Borgström (*C. A.* 12, 2507), on the remelting of algonite and whitneyite were checked by x-ray study. The results obtained confirm the view of B. that different products, namely Cu_2As and Cu, are formed. G. SCHWOCH.

Two minerals from the Belgian Congo. I. A non-pyro-electric tourmaline. II. A colorless chrysoberyl. J. MELOAN. *Bull. sci. acad. roy. Belg.* [5], 16, 996-1000 (1930).—The tourmaline occurs in fine needles, very pleochroic, $d = 3.0$, n about 1.636. The chrysoberyl occurs in hexagonal prisms. Measurements are given. R. S. DEAN.

Chemical and crystallographic investigations on lithidionite from Vesuvius. GUIDO CAROBBI. *Rend. accad. sci. Napoli* [3], 36, 21-31 (1930), cf. *C. A.* 25, 1462.—Lithidionite occurs both in vitreous and crystal forms. The crystal phase is richer in Cu, and samples contg. 13.29% CuO had sp. gr. 2.737-2.742, the vitreous phase, contg. 2 to 3% CuO had 2.483-2.373. Analyses of 3 samples: (a) vitreous, (b) vitreous plus crystal, and (c) crystal, are given. They establish its formula to be essentially $(\text{Cu}, \text{Na}, \text{K})_2\text{Si}_2\text{O}_7$, with CaO, MgO, Fe_2O_3 , and PbO present as minor constituents. The lithidionite probably contained $\text{Fe}_2\text{Si}_2\text{O}_7$ in solid soln. It was microscopically heterogeneous, and small quantities of wollastonite, and possibly of rivaite, were sepd. from it. It is monoclinic, optically—, $2V = 124^\circ$, the n_p 's are $n_p = 1.548$ and $n_m = 1.574$. Selected crystals contg. 13.29% CuO melted at about 780° to a blue glass which did not devitrify on cooling. An analysis of a sample of *neocianite* agrees well with that of lithidionite, establishing their identity. Arguments are advanced in support of the existence of *rivaite* as a definite mineral species. R. H. LOMBARD.

Genesis of Japanese acid clay. K. KOBAYASHI and K. YAMAMOTO. *Chem. News* 142, 116 (1931), cf. *C. A.* 24, 2352-3.—The acid clay is formed exclusively along an intrusion of liparite through pre-tertiary strata. Chalcedony, opal, pumice, primary kaolin and fine crystals of pyrite occur as impurities. Its formation is attributed to the decompn. of interposed Na feldspar and Na silicate. W. L. HILL.

Quinquennial review of the mineral production of India for the years 1924 to 1928.

Records Geol. Survey India 64 (1930). Jadeite. J. COGGIN BROWN 146-54.—The jadeite occurs as large lenses (5-7 ft thick) in albite intruded into partially serpentinized peridotites. Alluvial jadeite is also produced. Mica. CYRIL S. FOX 233-54.—Most of the Indian mica comes from pegmatites traversing mica schists. The pegmatites usually have a quartz core with marginal feldspar. Between these 2 are the largest mica books. Marketable mica is also found along the contact between dike and schist. The mica bearing pegmatites are recrystallized products of the schists. Less than 1% of the quarried rock is marketable mica. Other associated minerals are biotite, tourmaline, garnet, apatite, beryl and rarely samarskite, pitchblende and monazite. Ruby, sapphire and spinel. J. COGGIN BROWN 273-6.—The following gems are produced in Burma: quartz (amethyst, etc.), apatite, beryl (aquamarine), chrysoberyl, epidote, garnet, idioite, lapis lazuli, feldspar (moonstone), olivine, phenakite, tourmaline (rubellite), topaz, zircon. Most of the precious stones, and all the rubies and spinels, have been derived from bands of cryst. limestone in Archean gneisses. Zircon. W. A. K. CUKISTIS 312-3.—Zircon occurs (6%) in the Travancore sands, worked chiefly for ilmenite. Bauxite. C. S. FOX 323-43.—India contains large deposits of bauxite, largely unexploited. Uses are extensively discussed. Steatite. E. L. G. CLUGG 437-40.—Steatite is one of the most widely distributed minerals in India. A. H. E.

The Tadzhikistan phosphate region. R. M. ZBORIK. *Uchënyye i Urazhnye (Fertilizers and Crops)* 2, 204-16 (1930).—Z gives the geology of the region, petrography and mineralogical components of the rocks and of the phosphate deposits, their extent and approx. analyses. J. S. JORRE

Hydrazulite and sulfurous bauxite in Istria. T. KORMUS. *Bijele Kokske Lapke* 63, 269-77 (1930).—Bauxites of Santa Domenica and Portole show postvolcanic influences, e.g. a bluish gray bauxite with 10% S. Hydrazulites of the same region are probably of secondary origin formed independently of the bauxites, since they contain no Ti nor S, but have $H_2O > 30\%$. Several bauxite analyses are published. S. S. DE FÉLÉLY

Coal formation, mountain formation and bauxite formation in Hungary. E. VADÁZ. *Bijele Kokske Lapke* 63, 213-20 (1930).—The largest coal occurrences of Hungary were formed in post-orogenic sinking periods. Coal formation generally took place during the change of earth surface caused by the orogeny; the formation of bauxites, on the contrary, is the dry-land formation product of an ended orogeny. S. S. DE FÉLÉLY

The Kuréjka graphite occurrence of the Tunguzka coal basin, Siberia. A. PÉCELÉ. *Bijele Kokske Lapke* 63, 242-52, 277-81 (1930).—A coal basin covering about 900,000 sq. km. consists of 2 types of rock deposits. Coal of younger formation (Angaratype) is weakly metamorphic, contg. rather considerable ash. Coal of older formation (Jemsej type) is strongly metamorphic. The 14-m. amorphous graphite deposit of Kuréjka lies on a 1.5 m. lime layer. The proved coal content of the basins is estd. at 1,200,000 metric tons. S. S. DE FÉLÉLY

A new physicochemical explanation of the formation of humus, peat and coal. The significance of biological factors in these processes. J. ZOLCINSKI. *Wiss. Arch. Landw., Abt. A, Folge 4*, 193-228 (1930).—Decompn. of org. material of both plant and animal origin follows the same basic laws as are observed for the decompn. of rocks and minerals, in which simpler, more stable compds. are formed. Seventy references are given. W. GORDON ROSS

Comparison of the behavior to thermic action of rock from Ragusa and from Abruzzo. M. G. LEVI AND C. COLLINA. *Atti III congresso naz. chim. fra affluenti* 1930, 740-52.—The rocks studied were asphaltic calcites. They were subjected to fractional distn. to sep. liquid, solid and gas. A plant design is shown to sep. these substances commercially, a part of the rock being used for fuel. The residue can be used for CaO manu. E. M. SYMONS

Sapphires. J. W. HOWARD. *J. Chem. Education* 8, 613-24 (1931).—The sources, mining, sorting, grading and properties of natural sapphires are described. Details of the prepn. of synthetic sapphires are given and the methods of distinguishing natural and synthetic sapphires discussed. E. H.

Classification of rocks. F. PARR. *Bijele Kokske Lapke* 63, 536-63 (1930).—The classification of Hungarian rocks is discussed. S. S. DE FÉLÉLY

Astrophyllite-bearing nephelite syenite gneiss—found as a boulder in Kihitelyvaara, eastern Finland. PENTTI ESKOLA AND TH. G. SAHLSTEIN. *Bull. comm. geol. Finlande* 1930, No. 92, 77-88.—The mineralogical composition of the boulder was nephelite 34.27, plagioclase 27.92, K feldspar 17.15, amphibole 13.86, aegirine 0.14, astrophyllite 2.30, catapleite-like mineral 0.90, titanite 1.01, cancrinite 0.18, apatite

0.45%. It is accordingly highly alk and unlike any outcrops in the vicinity.

R. S. DEAN

The disintegration of rapakivi. PRATTI ESKOLA *Bull comm geol Finlande* 1930, No 92, 96-105.—The Finnish word rapakivi means crumbly stone and is used to designate certain granites which disintegrate spontaneously. The disintegration of rapakivi seems to presuppose the coincidence of several different circumstances. Its main condition seems to be a comparatively simple texture with rather smooth boundary surfaces between the mineral grains. The consistency of the rock may have been, and in some cases certainly has been, loosened by slight disturbances in the crust. Where layers thus predisposed become exposed to air and rain water and subjected to the influence of temp. changes, there the rock will break up into grit. At the same time chem. weathering sets in causing decompn. of the Fe rich lepidomelane and oxidation of its Fe⁺⁺ into rusty products. Part of the anorthite in the plagioclase, moreover, may be dissolved and some of its Al₂O₃ may remain in the insol. portion. From a rock which like the rapakivi is available in the form of loose grit the heavy minerals may be easily sepd. by means of panning. The rapakivi was thus found to contain considerable amts. of zircon and ilmenite, but no other minerals heavier than hornblende.

R. S. DEAN

Rocks of the upper Bargousin and Namama regions in Transbaikalia. PENTTI ESKOLA *Bull comm geol Finlande* 1930, No 92, 100-40.—The following rocks occurring in the Bargousin region are described: graphite bearing limestone, amphibolites and gneisses, quartzite, granite. In the Namama region, cryst. limestone and dolomite, quartzite shale, greenschists and metamorphic andesites are described as super-crystal and quartz keratophyre, orthoclase porphyry, granite, granodiorite, hornblende gabbro and hornblende are described as intracrystal. Cu ore also occurs in the cryst. limestone bounded on the east by quartz keratophyre and on the west by Namama diorite. The chem. nature of the igneous rocks of the Bargousin magmatic province is discussed in detail.

R. S. DEAN

The eruption of the volcano of the Kaménis (Santorini) in 1925. CONST. A. KTEAS. *Bull volcan* 4, 7-49 (1927), cf. *C. A.* 20, 564, 21, 840, 24, 2080.—A general description is given of the eruption. Two analyses of fumarole gases and 4 of lavas are given.

H. S. WASHINGTON

Report on studies on the last eruption of the Kaménis (Santorini). CONST. A. KTEAS. *Bull volcan* 4, 171-6 (1927).—An addn. is made to previous papers on the eruption without new analyses.

H. S. WASHINGTON

Report on the work of the Laboratory of Petrology of the University of Athens, relating to the study of the volcanoes of the Aegean Sea. CONST. A. KTEAS. *Bull volcan* 4, 182-7 (1927) (map).—Brief notes are given on the petrography of some little-known Greek island volcanoes. The lavas are mostly dacites and andesites. H. S. W.

The basaltic volcanoes of southern Indochina. M. F. BLOUDEL. *Bull volcan* 4, 193-4 (1927).—Very brief notes are given. No analyses are included.

H. S. W.

The chemico-mineralogic characters of the Tertiary intrusive and volcanic rocks of North Africa. A. LACROIX. *Bull volcan* 4, 199-204 (1927).—A summary paper. Although no analyses are given, the rocks are classified according to the C. I. P. W. system.

H. S. WASHINGTON

First observations on the mineralogic and chemical composition of the Mesozoic and Tertiary lavas of eastern China. A. LACROIX. *Bull volcan* 4, 205-17 (1927).—Brief petrographical descriptions are given with 20 chem. analyses.

H. S. WASHINGTON

The lithologic constitution of the South Central Pacific. A. LACROIX. *Bull volcan* 4, 218-31 (1927).—A general petrological description is given, no analyses are included.

H. S. WASHINGTON

Collection of chemical analyses of Russian eruptive and metamorphic rocks. Z. NÉMOVA AND F. LOEWINSON-LESSING. *Mem Com Geol* (Moscow), New series, No 186, 361 (1930).—The collection numbers 1676 analyses. The text is in Russian, except that rock names are also given in French.

H. S. WASHINGTON

Unusual depositional stalactites in a lava-tunnel at Mount Albert, Auckland. J. A. BARTRUM. *New Zealand J. Sci. Tech.* 12, 188-92 (1930).—The material appears to consist of crust like layers of small white flakes of opal intermixed with kaolin and sepd. irregularly by similar discontinuous thin layers of granular white kaolin.

H. C. PARISH

Shippings of clay territories at Budapest. A. VÉNDL. *Magyar Mérnök Épités segyelt. Közönye* 64, 65-72 (1930).—Shippings are caused by a thin water-contg. layer within the loess covering Kiscelli clay strata.

S. S. DE FINÁLY

Translation gliding in crystals of the NaCl structural type (BLUGGER) 2. Membranes of spores and pollen. IV Fossil sporopollenins from Tasmanite and Russian lignite (ZETSCHE, *et al.*) 11D.

FARNHAM, C. MASON Determination of the Opaque Minerals. New York: McGraw-Hill Book Co., Inc. 236 pp. \$3.50

NANNI, RAFFAELLO I soffioni e i laghi della Toscana e la industria boracifera. Rome Reviewed in *Analyst* 55, 215(1931).

9—METALLURGY AND METALLOGRAPHY

D. J. DEMAREST, H. W. GILFILL AND RICHARD KIMBACH

Development and present state of metallurgy. F. ALVÉDER *Stands Kells Lapok* 62, Supplement 3-4(1929) —A historical and descriptive treatise including a bibliography. S. S. DE FINLEY

Origin of flotation. L. MICHEL *Rev. esp. miner y metal.* (Santiago) 5, 702-4(1930) —A historical review. E. I. S.

Properties of the bog iron ore from Hokkaido and of the brown iron ore from Chohsen as determined by thermal analysis. KATSUMI INOUE, *Toku-to-Hagetsu* (J. Iron and Steel Inst. Japan) 17, 99 100(1931) —From thermal analysis of these ores, the heat necessary to expel the combined water was calculated. The amt. of heat required for the bog iron ore was larger than that for the brown ore at the same rate of heating, though the decompn. temp. of the former was lower than that of the latter. It attributed this to the internal strain in the grains of the Chohsen brown ore. The sintering temp. of the brown ore of Chohsen was higher. M. KURODA

Investigation on the reduction processes of iron ore. III. BERTIL STALLHANS AND TORE MALMBERG. *Jernkont. Års.* 114, 69-72(1930) —The reduction of Kiruna magnetite in H_2 and H_2 -CO mixts. has been investigated. At 930° 0.2 mm. grains were completely reduced by H_2 in 2 min., 2.0 mm. grains in 7 min. and 4.0 mm. grains in 17 min. The rate falls off with the temp. but there is a marked discontinuity at 910°, at which temp. the time required is 2.5 times that at 890°. Addn. of CO to the H_2 slows up the rate quite markedly, the rate in 100% H_2 being 7 times as great as in 100% CO. Pure CO also gives the same discontinuity as H_2 at 910°. Mixts. show the same but less markedly. The temp. at which the reduction becomes exceedingly slow, 930°, corresponds to the transition point of α -Fe to γ -Fe. H. C. DUFF

Statistics on the Swedish iron industry during recent years. ARVID JOHANSSON. *Jernkont. Års.* 114, 623-34(1930) —Consumption of charcoal decreased from 2.72 million cu. m. in 1924 to 1.98 million cu. m. in 1929. The production of Fe ore remained essentially const. at 330,000 to 670,000 tons. The value of the export was 8.5% of the total value of all exports, and the value of imports 4.9% of the total. H. C. DUFF

Superroasting of zinc blende by different methods. E. PROST. *Rev. min. Suisse* 72, 234-40, 269-77(1929) E. I. S.

Free energy of some copper compounds. MERLE RANDALL, RALPH F. NIELSEN AND GEO. H. WILST. *Ind. Eng. Chem.* 23, 368-400(1931) —A review, intended primarily to aid Cu metallurgists, of the equilibria in various reactions of Cu compds. The free energies of various Cu compds. important in the metallurgy of Cu are given. References (102) to articles in the literature used in obtaining the data are given. J. B.

Sand tests in the foundry. J. HICK. *Foundry Trade J.* 44, 191-3, 197(1931). —Methods for making sand tests and their application to foundry control are described for both facing and core sands. Data sheets are included showing various phys. properties of many types of sands found in several of Scotland districts in England. D. S.

Investigation of furnaces used in rolling mills. L. A. RICHTER. *Stat. u. Eisen* 51, 377-83(1931) —A review of the operating practice of the furnaces used in rolling mills in Austria. The type of furnaces, the method of operation, the heat balance, the oxidation losses, the dependence of the fuel consumption on the type of material rolled and the ratio of the hearth surface to the total furnace surface are discussed. J. A. SILLARD

Rotary-hearth furnace for heat treating. R. E. BARKER. *Heat Treating and Forging*, 16, 119-90(1930) —A rotating hearth mounted on ball bearings in oil and motor-driven is described. It burns producer gas, natural gas, oil or tar, is adaptable to all classes of steel and has proved ideal for "open" annealing of all types of material. J. BALOGHIAN

Utilization of slag from blast furnaces in Japan. TAIROH KURODA. *Tetsu-to-Hagane* (J. Iron and Steel Inst. Japan) 17, 118-22(1931).—Statistics of the production of pig iron and slag in Japan are given. The main uses for slag are in the production of cement, slag ballast brick and slag wool. The analyses of slag and its uses in 8 main blast furnace works in Japan are given. M. KURODA

The slag in the basic open-hearth steel furnace. R. BACK. *Stahl u. Eisen* 51, 317-24 351-60(1931).—In 19 melt tests carried out in the basic Siemens Martin furnace the basicity of the slag as characterized by the CaO/SiO_2 ratio was followed. During the melt-down period, with increasing basicity of the slag, the amt. of Fe in the form of oxides increased, while the SiO_2 content decreased. No definite relationship could be found between the Mn content of the slag and the basicity during the melt-down period. Increasing amts. of Mn addns. caused increased Mn content in the slag during this period. During the boiling period, which, in the scrap-Fe process, is the time between the end of the melt-down and the addn. of the finishing flux, the Fe and Mn contents increased with the increasing basicity of the slag. Five different types of slag classified according to the CaO/SiO_2 ratio, were investigated: acid slag with a ratio of 1.3-1.5, medium basic 1.5-1.8, good normal 1.8-2.0, basic 2.0-2.4 and highly basic over 2.4. The above ratios refer to the compn. of the slag at the end of the boiling process. For the regeneration of Mn from the slag the acid slags, contrary to expectation, gave very good results. The best results were obtained with the medium basic slags. Even with the good normal slag the Mn had the tendency to go back from the slag into the bath. In the basic slag the tendency for the reduction of Mn was found to be very slight. The highly basic slag gave the poorest results. At very high temps. the results with highly basic slag were somewhat better. The Fe content of the slag during the boiling period, as in the melt-down period, is a direct function of the basicity of the slag. The Mn content is governed by the Fe content: with increasing basicity more FeO and Fe ferriite are formed in the slag and these have an oxidizing effect on the Mn in the bath. It is emphasized that the basicity of the slag has a detg. role in the transformation of Fe and Mn. The slag test is described; it gives a good indication of the compn. of the slag and is of importance in cases where the P and Si contents of the raw materials are subject to great variations. J. A. SZILARD

Method of slag determination. H. KJERMAN. *Jernkontorlets Ann* 1929, No. 4, 181-99.—A discussion of (1) the slag of different kinds in steel, (2) its microscopical detn., (3) the relation between slag content and strength and (4) the sulfate content in Cr ball bearing steel. E. I. S.

Some modern tendencies in siderurgy and the establishment of this industry in the Argentine Republic. SYLV WASSERMAN. *Anales soc. cienc. argentina* 111, 65-92(1931). E. H.

The preparation of the raw materials at the Röschlingen Iron and Steel Works in Völklingen (Germany). A. WAGNER. *Stahl u. Eisen* 51, 217-25(1931).—The parts of the plant for the prepn. of the raw materials including breaking, screening and sintering are described and their economies discussed. J. A. SZILARD

Choice of raw materials for malleable cast iron. I. II. III. J. V. MURRAY. *Metallurgia* 1, No. 3, 107-9; No. 4, 161-2, 165, 2; No. 9, 101-2(1930). E. J. C.

Statistical evaluation of pig-iron analyses. K. DAERES. *Stahl u. Eisen* 51, 202-4(1931).—A method is shown for the evaluation of the daily analyses of 3 blast furnaces and one mixer during 3 months. The deviation from the av. values is greater for the blast furnaces than for the mixer. Curves are given which show the relationship between the S, Mn, Si and P contents. J. A. SZILARD

Advantages of fine-grain over coarse-grain pig iron for production of high-temperature iron. R. STORZ. *Z. ges. Gesesserspraxis* 51, No. 3, 25-8(1930), cf. *C. A.* 24, 42.—S seeks to show that 2 pig irons with apparently the same chem. compn. do not necessarily possess the same phys. characteristics. E. I. S.

The manufacture of cast-iron rolls. E. PEPPER. *Stahl u. Eisen* 51, 345-51(1931).—The manufacture of loam and chilled cast rolls is described, and also the development of alloyed rolls. J. A. SZILARD

The development of the research program of the steel-making processes on a physicochemical basis. H. SCHENCK. *Stahl u. Eisen* 51, 197-202(1931).—In the effort to reduce production costs and improve the quality of steel products, comparatively little attention was given to the basic chem. reactions involved in steel making. The unsolved problems of steel making are listed and a research plan is developed which would make coöperation between steel plants, exptl. stations and scientific research institutions possible, to obtain a better knowledge of the metallurgical processes. J. A. SZILARD

Refining of aluminum and its alloys by treatment with chlorine gas and nitrogen. W. KOCN *Metallwirtschaft* 10, 69-72, 85-8(1931).—Lab. expts were carried out to det. the influence of impurities in Al upon the temp. of chlorination. The samples were heated in a Cl stream in combustion boats, the temp. of the beginning of volatilization was detd. and the volatile matter analyzed. The chlorination should conform with the following theoretical considerations: (1) It is a function of pressure and temp. and subject to the mass law. (2) The chlorination temp. of a pure metal must be reduced by alloying it with another constituent of a lower chlorination temp. The bigger the difference between the chlorination temp. of the base metal and of the alloying constituent, the greater will be the influence of such an addn. (3) The volatility of the chlorides formed is an essential factor in the speed of chlorination. It is increased by the presence of compds. of a lower temp. of volatilization. Impurities act proportionally to the difference between the volatility of their chlorides and those of the base metal. The expts. brought out that the temp. of volatilization of pure Al is reduced by impurities, especially Fe and Si. At the same time the volatility of $AlCl_3$ is being increased. The impurities also interfere with the formation of a protective film, Si to a greater extent than Fe. They form centers of corrosion, thereby rendering the metal less compact and increasing the surface, both of which factors interfere with the film formation and facilitate the attack. Refining tests were made on a technical scale by introducing Cl into the molten metals and finally removing it by blowing N through the melt. Various types of com. Al and of Al-Fe, Al-Si and Al-Cr alloys were thus treated. In alloys contg. 0.8, 1.6 and 1.86% Fe, the chlorination resulted in a fine and dense grain, while the original material was coarsely cryst. The plasticity and toughness of the alloys was also increased. No change in the chem. compn. of the alloys could be ascertained. An alloy contg. 3.52% Fe and 0.43% Si showed the grain refining effect of the chlorination especially distinctly. In this case the treatment decreased the Fe by 0.37%. In an alloy contg. 11.7% Fe and 0.45% Si no refining effect of the chlorination could be found, except the removal of considerable quantities of gases. The Fe was reduced by 0.67%. In Al-Si alloys contg. 0.8, 1.3 and 2.1% Si the chlorination had a dispersing effect upon the segregations found in these alloys, which consist mainly of Al-Fe-Si compd. The chlorination of silumin had a slightly coarsening effect upon the metal grain, because of the chlorination and volatilization of the colloidal Na particles. An Al-Cr alloy contg. 2.8% Cr showed only a slight refinement of the Al-Cr segregation. An alloy contg. 1.9% Cr and 1.4% Fe showed a much more distinct improvement. The eutectic and the Al-Cr segregation were finely divided and the Al grain was considerably refined. Expts. made with com. grades of Al and Al scrap showed in each case that the chlorination had the effect of refining the grain and dispersing the impurities. The chlorination has a favorable effect upon the mech. properties of the treated materials. The ductility and the bending properties are improved and the hardness is increased. The main field of application of the Cl and N treatment will probably lie in the refining of Al scrap. While a selective volatilization of the impurities appears to be impossible, a no. of valuable phys. improvements are obtained, such as: removal of gases, clean grain boundaries, grain refinement and homogenization, and an increase of fluidity. The castings obtained show a dense structure, free of blow holes, and a smooth, shiny surface. Most of the mech. properties show distinct improvements.

LEOPOLD FESSEL

An investigation of core oils. CARL H. CASBERG and CARL E. SCHUBERT. Univ. Ill. Eng. Expt. Sta., *Bull.* No. 221, 5-22(1931).—The oils investigated were those commonly used in core work, i. e., the drying oils linseed, China wood and perilla, together with the semi-drying and non-drying such as soy-bean and kerosene. The object was (1) to det. any possible relation between tensile strength of the cores and the chem. and phys. properties of the oils and (2) to note the effect of moisture on tensile strength. Tests made on the oils included ash content, sp. gr., flash and fire points, sapon. and I nos. Results showed that no definite relations exist between these properties and the av. original tensile strengths of the finished cores, although the initial tensile strength was found to be proportional to both the sapon. and I nos.

H. L. OLIN

Progress in metallurgical research. C. H. DESCH. *Foundry Trade J.* 44, 105-6 (1931).—The impregnation of solid ores by C must play an important part in the reactions of the blast furnace. Ore so impregnated with C is readily reduced at a higher temp., the fine solid C in the mass being a much more powerful reducing agent at 750° than is CO. The solid C may also form a somewhat impervious layer around the ore granules and hinder the reduction. The exact conditions detg. the form that the C will take have yet to be learned. Endurance tests, rather than short-period tensile

tests, are necessary at high temps on metals, as creep under prolonged stress bears no const relation to the mech properties under static conditions. An abbreviated test, described by W H Hatfield has been found useful as an opprox guide. That stress is found which at the given temp, produces on elongation not exceeding 0.5% in the 1st 24 hrs., and produces no further change of length in the next 48 hrs., when measurements are made such as would detect an elongation of the order of 1 millionth in. per in. per hr. The safe stress is then taken as $\frac{1}{2}$ of this limit. Both magnetic aging and strain etching are dependent upon the sepn of nitride from solid soln. There is no direct relation between them. X ray analysis has greatly assisted in carrying out an immense amt of research work relating to changes which metals undergo during deformation. App has been devised for measuring pressures developed at the rolls and at the dies and records have been kept of power consumption with different degrees of reduction. By quenching in a blast of H₂ and by using the string galvanometer cooling curves have been taken during the short period of quenching. The evidence thus obtained indicates that martensite is a heterogeneous mixt. contg finely divided ferrite and cementite and not a supersatd soln of carbide in α iron. H C. P.

Literature on the use of the x-ray. II. VICTOR S POLANSKY, *Heat Treating and Forging* 16, 1011-14 (1930), cf C A 24, 5690.—A bibliography is given of the literature available in the Carnegie Library of Pittsburgh on x ray examn of ferrous metals and alloys for authors names from F to K, inclusive. The list has been brought up to May, 1930. J BALOZIAN

The deformation of metals under prolonged loading. I. Flow and fracture of aluminum. D HANSON AND M A WHARTON, *J Inst Metals*, Advance copy No 554, 29 pp (1931).—Al samples were subjected to static loads at room temps. and at 250°. The mode of deformation was detd by microscopic examn of the polished and unetched surface of the test specimens. Samples consisting of various sized crystals including single crystals, were used. Conclusions: "(1) The extension under a prolonged load that will ultimately break the metal may be conveniently considered as consisting of 3 periods: (a) a period of primary extension, during which the rate of flow diminishes, (b) a period during which flow is very slow or even suspended, (c) a period during which the extension again continuously increases until fracture occurs. (2) During the first period the flow of the metal is the result of slip within the crystals, and flow diminishes because of the hardening effect of slip. (3) The behavior during the second and third periods varies with different specimens. (4) Three types of failure have been recognized under creep conditions: (a) failure by intercryst cracking: cracks form while the elongation is low, and the metal fails with little reduction of area at the fracture. (b) failure by the resumption of slipping within the original crystals, leading to a fracture of the normal type. (c) failure through the recrystn of the metal, removing strain hardening and permitting further flow of the softened metal, the metal pulls out to a point fracture, with a high elongation. (5) Al can fail to "creep" at room temp. (6) Al consisting of uniform aggregates of crystals fails at 250° by intercryst. cracking. (7) Al aggregates at room temp., and Al single crystal specimens at all temps., fail under creep conditions by shear along the slip planes. (8) In single crystals when flow recommences slip occurs only on some of the surfaces that had previously slipped during the primary flow. (9) The prolonged action of a suitable stress can cause disintegration of a metal either at the junctions of the grains or on the surfaces of previous slip, and failure (i. e., fracture) must be regarded as having commenced when this weakening occurs. (10) Under suitable conditions, e. g., in crystal aggregates with moderate rates of flow at 250°, or with very slow rates of flow in single crystals, slip is distributed over a very large no. of slip surfaces, and a polished surface of the metal may show little or no sign of slip bands. (11) No change in density is believed to occur when a crystal is distorted by moderate amts.: changes in density in crystal aggregates are thought to occur at grain boundaries. (12) The view is expressed that rupture of the crystals commences along slip planes formed at an early stage in the deformation of the metal. (13) The results are consistent with the theory of slip and rupture previously advanced by Gough, Hanson and Wright." Four references are included. J L GRIGG

The mode of deformation of a single crystal of silver. H J GOUGH AND H L COX, *J Inst Metals*, Advance copy No 552, 18 pp (1931).—A single crystal of Ag was subjected to alternating torsional stresses and the surface examd to det the mechanism of deformation. Numerous slip bands were formed on the octahedral planes, but there was no definite evidence of twinning even when the stresses were applied until failure. The slip bands might have been narrow twins, but this appears to be improbable. No twins could be produced by compression. Because of the ease with which twinning

occurs in an aggregate of small crystals, the mechanism of deformation in a single crystal must be essentially different from that in the aggregate. J. L. GREGO

Microstructures of fifteen silver Greek coins (500-300 B. C.) and some forgeries. C. F. FLAM. *J. Inst. Metals*, Advance copy No 550, 10 pp (1931)—All of the genuine coins had been struck, while only one of the forgeries showed any indication of striking. Some of the coins had apparently been struck hot. Treatment differed at different mints. Analyses indicate that Cu had been intentionally added. J. L. GREGO

New uses for aluminum. L. JAKOVY. *Bányás Kohás Lapok* 63, 236-41 (1930)—A review. S. S. DE FINÁLY

Some properties of metallic cadmium. C. H. M. JENKINS AND G. D. PRESTON. *J. Inst. Metals*, Advance copy No 556, 32 pp (1931)—Hardness and tensile values of rolled Cd show that the temp. of working has a pronounced effect on these properties. In ordinary tests the worked Cd appeared to be stronger than the cast, but the conditions were reversed in the long-time ball test. Material in the cold rolled condition was completely recrystallized. In such material appreciable grain growth was found after 5 years. No evidence was found of an allotropic transformation. By x-ray examn. the parameter of the unit cell was found to be $a = 2.9721 \pm 0.0005$ A. U., with an axial ratio of 1.8854 ± 0.0005 . In cold rolled Cd the (1011) plane tends to lie parallel to the surface, and the normals to the basal plane (0001) are inclined at about 65° to the normal to the surface of the sheet. In hot rolled sheet another preferred orientation is found. Seventeen references are given. J. L. GREGO

Three crystalline modifications of electrolytic chromium. KUMAZO SASAKI AND SINKITI SEKITO. *Trans. Electrochem. Soc.* 59 (preprint) 7 pp (1931).—See C. A. 25, 1195. C. G. F.

The independence of the hardness and the hydrogen content of electrolytic metals. GUICHARD, CLAUSMANN, BILLON AND LANTHONY. *Compt. rend.* 192, 623-5 (1931)—The exceptional hardness of electrolytic Fe, long noted, has been attributed, on the one hand, to the fine state of diffusion of the metal and, on the other, to the presence of a compd. of H and Fe having a characteristic metallographic structure. The authors have repeated the expts. of Hugues (C. A. 20, 572) and reached an entirely different conclusion. The results of their expts. indicate that it is necessary to exercise considerable caution before attributing the remarkable mech. properties of electrolytic metals to the presence of occluded gases. O. W. ELLIS

Notched-bar impact tests. E. HÖGA. *1st Communications New Intern. Assoc. for Testing Materials* 1930A, 213-6.—Notched bar impact tests of heat-treated steel contg. 0.15% C showed the effect of the superheating after heating for 1, 2 or 3 hrs. at 1100°. Examn. of plates from an exploded boiler showed the energy absorbed in fracture of notched bar test pieces after various heat treatments and at different temps. of the test bar during the impact test. E. M. SYMONS

Cyaniding and salt-bath working. J. W. URQUHART. *Heat Treating and Forging* 16, 989-93 (1930)—The details involved in successful salt-bath hardening, and general heat treatment in salt baths are briefly described. J. BALOZIAN

Heat treatment of non-ferrous metals. A. H. VAUGHN. *Heat Treating and Forging* 16, 1036-8, 1041 (1930)—A brief survey is given of the present practice. J. BALOZIAN

Heat treatment of aircraft engine parts. R. R. MOORE. *Fuels & Furnaces* 9, 279-89 (1931)—The importance of proper heat treatment is outlined and methods of heat-treating steel and Al alloy parts are described. The chem. compn. and phys. properties of the Al alloys are tabulated. MORRIS SCHRERO

Variations in hardness of metals and alloys resulting from cold working. GUTCHARD CLAUSMANN AND BILLON. *Bull. soc. chim.* [4], 49, 173-85 (1931).—See C. A. 24, 4749. G. G.

Some phases of heat treatment of cylinder and alloy irons. F. J. WALLS AND A. HARTWELL, JR. *Trans. & Bull. Am. Foundrymen's Assoc.* 2, No 3, 805-96 (1931).—Tests were run to det. the relief of stresses in cylinder irons on heating to temps. of 750° to 1150°F. The hardness of the irons as affected by air-, water- and oil-quenching was investigated. Max. relief of stresses is obtained with min. growth and change in hardness by heating to 950°F., followed by slow cooling. Similar tests were made on Cr-Mo and Cr-Ni-Mn cast irons. C. H. LORIG

Wear resistance of nitrided nitralloy. V. O. HOMERBERG AND J. P. WALSTEAD. *Metal Progress* 18, No 6, 68-71 (1930)—A study of the wear resistance of nitrided nitralloy against itself, gray and white cast Fe, Monel metal and 3 bronzes shows com. possibilities of nitrided materials in engine or machine parts which are subject to severe wear conditions and which are difficult to fabricate. W. A. MUDGE

Tests on the resistance of cast iron to wear. R. SPATIER. *1st Communications Soc. Intern. Assoc. for Testing of Materials* 1930A, 64-73 (in French) — There is no standard method in use for measuring resistance to wear of cast Fe, and no basis of comparison of results obtained by various methods of testing this property. Tests were made by pressing at 25 kg. pieces of cast Fe shaped like a brake shoe 1 cm. thick and 3 cm. deep against a steel disk 4 cm. in diam. and 1 cm. thick, revolving on an axle at 156 r.p.m. tangential velocity 0.375 m. per sec. Phys. tests of the samples were made previously. It was impossible to recognize any relation between hardness, resistance to traction, resistance to flexure and resistance to shock. In general, wear was inversely proportional to hardness. After a certain period of testing, due to the grinding action of abraded material, less hard cast Fe often showed less wear than harder samples. Wear is faster at first than near the end of the test. Wear of cast Fe is proportional to that of the steel disk. Wear is greater in a continuous than in an intermittent test. Wear increases with increase in graphitic C, and decreases with increasing P content. E. M. SPOONER

Comparison of the physical properties of different sections of cast iron and of the standard arbitration test bar. M. V. HEALEY. *Proc. Am. Soc. Testing Materials* 30, Pt. 1, 299-324 (1930) — A hollow box, cast in green sand with a dry sand core was made with one wall $\frac{1}{2}$ " thick, one 2" and one nominally $\frac{1}{2}$ ", but actually, from the flexure test table, $\frac{1}{2}$ ". Standard 1.3"-diam. arbitration bars were cast from the same metal, in dry sand. In the different sections total C ranged from 2.74 to 3.05% combined, from 0.49 to 0.63% Si from 1.95 to 2.19%, Mn 0.55%; P 0.37%. In flexure tests on 12" supports, with bars 1" wide by 4, 2, 1 and $\frac{1}{2}$ " with 1" dimension vertical, deflection ran from 0.155 to 0.19" on the specimens from the casting, with no marked effect of size, while the arbitration bars gave 0.12". The modulus of rupture of the $\frac{1}{2}$ " thick section was 45,000 lb. and for the thinner sections varied from 45,000 to 52,500 with no marked effect of size. Averages showed a decrease with the size of the section but individual figures for each size scattered over about the same range. The arbitration bars, however, gave 63,000. Tensile tests showed 32,000 for the arbitration bar and 23,500-30,000 for the castings with increase in strength as the section diminished. The $\frac{1}{2}$ " section was a trifle out of line, probably through insufficient feeding. Brinell tests showed 193 for the arbitration bar and 150-170 for the casting. Brinells showed fair correlation with strength. Micrographs are shown. J. W. GILLET

Comparative tests on cast iron in Czechoslovakia. P. PRER. *1st Communications Soc. Intern. Assoc. for Testing of Materials* 1930A, 25-34 (in French) — Strength tests and cutting tests are described. E. M. SPOONER

Nitrogen in technical iron. III. The separation of nitrogen and carbon from α -iron as an example of the decomposition of a doubly supersaturated solid solution. WERNER KOSTER. *Arch. Eisenhüttenw.* 4, 140-52 (1930) cf. C. A. 25, 671 and following abstr. — The tempering isotherms were determined for a set of steels of varying N content for different quantities of dissolved C at 100° and the increase of the coercive force measured during the first rise. For any percentage of C there exists a linear relation between the total amt. of N and the coercive force. The isothermal sepn. of N from α -Fe supersatd. with N at 100° is interfered with if C in solid soln. is present at the same time. The quantities of N eliminated are smaller the higher the concn. of the dissolved C, and no sepn. takes place at all if a certain limit of C is exceeded. These limiting values can be plotted as a curve in the Fe-C-N system. The points along this curve indicate the amt. of N remaining in soln., these values depend on the amt. of C in soln. and are independent of the total amt. of N. The expts. give an example of the influence exerted by a constituent in a solid supersatd. soln. upon the sepn. of another. L. P.

Nitrogen in technical iron. IV. The combined action of cold working and nitrogen separation on the magnetic properties of technical iron. WERNER KOSTER. *Arch. Eisenhüttenw.* 4, 289-94 (1930), cf. preceding abstr. — The influence of various heat treatments on the magnetic properties of cold worked steels (0.012% N) is investigated. The coercive force is detd. of specimens (1) which have been cold worked (elongated 0-20%) and subsequently annealed at 100° up to 14 days, and (2) which have been treated as in 1 and then re-annealed for 1 hr. at temps. up to 700°. These expts. show that the coercive force is a result of (1) that attributable to cold working (increasing in this case with increasing elongation), and (2) that caused by N sepn. (decreasing here approx. linearly with the elongation). The relation between the coercive force and the degree of straining is independent of the order of treatment. The change in the magnetic properties of cold-worked steel between 100° and 300° is the result of the sepn. and soln. of N. The observations in the literature are explained on the basis of the results of these expts. J. BALOTIAN

The development of quality thin sheet iron. H KLEIN. *Stahl u. Eisen* 51, 189-96(1931).—A discussion is given of the increased demand for quality thin sheet iron, leading to improvements in the equipment of rolling mills, in the annealing and finishing of sheet iron. The methods of galvanizing and tinning and the properties of rust resisting steels are also discussed. J A SZILARD

Progress of the knowledge regarding slag (solid nonmetallic) inclusions in iron and steel. C BENEDICKS and H LÖRQUIST. *1st Communications New Intern Assoc. for Testing of Materials* 1930A, 345-7.—A review with bibliography of the more important progress since the author's 1927 paper (C A 23, 75). E J C

Continuous load experiments on various construction steels at high temperatures. ERNST POHL, HANS SCHOLZ and HUBERT JURETZEK. *Arch Eisenhüttenw* 4, 103-11 (1930).—Normalized flat test pieces of low-C steel (0.1% C), Mo steel (0.15% C, 0.34% Mo) and Ni steel (0.28% C, 2% Ni) were subjected at const. temp., at intervals of 50° between 300° and 500°, to definite loads, successively increased. The max extension reached at each load was recorded as a function of the time, and from the data obtained, curves were plotted giving the relation between load and elastic strain and between load and total extension. Within the range of temps. investigated the Mo steel proved to be most resistant to permanent flow under the loads applied. At 350° under a load of 175 kg./sq mm applied for 504 hrs. the Mo steel had not yet attained the condition of plastic flow. The elongation at rupture was 0.172%. H S VAN KLOOSTER

Steel hardening. H ESSER and W EILENDER. *Arch Eisenhüttenw* 4, 113-44, *Stahl u. Eisen* 50, 1616-7(1930).—After having proved in preliminary quenching tests that the heat effects occurring during the quenching of steel specimens in water cannot be detd accurately, the quenching was performed by means of gas. The heating and quenching of the samples were performed in a vacuum in a container designed for this special purpose. The specimens used were either thin sheets (0.5 to 1 mm. thick, 1-2 mm wide) or small balls 1-2 mm in diam. Quenching curves with varying cooling velocity were taken on C steels from 0.01 to 1.75%. The results permit certain conclusions to be drawn as to the nature of Ar' and Ar''. The heat effect occurring between 300° and 400° with cooling velocities of about 100° to 200°/sec. is assumed to be a tempering effect due to the essentially decreased cooling velocity within this temp range. In considering these results, the authors arrive at a new explanation for the causes of temper brittleness. A hardening diagram for a cooling velocity of 1000°/sec is designed, indicating that the heat effects during quenching are similar to the transformations of the Fe-C diagram. It must probably be assumed that martensite is not a pseudo solid soln but a heterogeneous submicroscopical mixt. of α -iron and cementite. It is furthermore concluded that all Fe-C alloys, even those with very low C contents, are hardenable if they are sufficiently rapidly cooled. The microscopical investigations are in accordance with the results of the thermal analysis. The development of a new theory of hardness is attempted, which explains the hardness in quenching the austenite with crit. velocity by the following 2 factors: (1) the high degree of dispersion of the cementite; and (2) the pronounced distortion of the α lattice during the γ - α transformation and the cementite pptn in the temp range of a decreased plasticity of the steel. Both factors account for the hardness. G G NEUENDORFF

Stainless steels used in heavy machining. CHARLES E. HERD. *Metal Progress* 19, No 1, 44-9(1931).—Special applications in paper mill machinery and hydraulic turbines are described. W. A. MUDGE

Steels used in Ford industries. J. L. McCORMAN. *Metal Progress* 19, No. 3, 32-9 (1931).—A general review. W. A. MUDGE

Steel and its heat treatment for parts that must resist wear. H W. MCQUAIN. *Heat Treating and Forging* 16, 1159-62(1930), cf. C. A. 25, 2090.—The different steels and treatments which can be used for obtaining hard surfaces are enumerated, discussed, and classified according to their relative cost. J BALOZIAN

Creep of steel at elevated temperatures. P. G. McVETTY. *Mech. Eng.* 53, 197-200(1931).—The principles underlying the creep of steel at elevated temps are discussed. The phenomenon of creep is defined, typical creep curves are shown and interpreted, and the metallurgical significance of creep phenomena and its significance to the designer are pointed out. G G NEUENDORFF

Some long-time tension tests of steels at elevated temperatures. J. J. KANTER and L. W. SPRING. *Proc. Am. Soc. Testing Materials* 30, Pt I, 110-32(1930).—Very detailed evidence is given to show the value of a logarithmic method of plotting for comparing strain hardening effects and to bring out the fact that rates of creep are not necessarily const. Long-continued tests often show a decreasing rate because of strain-hardening.

Comparison of short time tensile and long time flow tests shows that at the higher temps the short time tests do not reliably reflect the properties. At the higher temperatures large grain size whether obtained in cast as compared with wrought alloys, or by annealing as compared with quenching and drawing or with rolling without annealing, tends to reduce creep. Step loading running awhile at one load and then increasing the load on the same specimen may give misleading results. The discussion deals primarily with methods but the curves give much actual data on the flow properties of carbon and Ni-Cr cast steels and on wrought carbon, Cr-Mo, Ni-Cr, stainless and 18-8 Cr-Ni steels. Some of the curves show results from observation of the same specimen for a year or more.

H. W. GILBERT

Structural changes in annealed soft steel. W. HIFKE AND W. BAENSCHIEDT. *Arch. Eisenhüttenw.* 4, 91 (1930) (1930).—Mild steel (0.16% C, 0.02% Si, 0.35% Mn, 0.005% P and 0.007% S) was cold rolled to different degrees (as far as 60% reduction in area) and annealed in limestone, in air and in a lead bath for periods up to 100 hrs. at temps between 700° and 800°. Large columnar grains appear on the periphery at a degree of deformation around 1%. The phenomenon disappears when the reduction in cross sectional area reaches 6%. The columnar growth can be used as an indication of previous cold work deformation in conjunction with Fry's method of etching. Grain growth occurs also in the interior but uniformly, as soon as the C has diffused towards the outside. The loss of C, even in an annealing bath of lead, is about 40%. Prolonged annealing slightly below A_1 in an oxidizing atm. at degrees of deformation above 30% produces unusually large grains which originate in spots that are rich in P. When the annealing process is carried out between A_1 and A_2 , the C is displaced from the spots that contain a considerable amt. of P and grain growth occurs as a result of the diffusion of P.

H. S. VAN KLOOSTRA

Thermomdiffusion of elements in steel. H. JOHN H. HAUSKA. *Heat Treating and Forging* 16, 1530-3 (1930) cf C 125, 1230.—The thermomdiffusion of C in carburizing and cyaniding and of O, and of decarburization as a thermomdiffusion phenomenon are discussed.

J. BALOZIAN

Nitrided steel and its properties. J. H. HIGGINS. *Heat Treating and Forging* 16, 1527-9 (1930).—The existing practical conditions confronting the com. nitriders are discussed.

J. BALOZIAN

German versus American sheet steel. EDWARD S. LAWRENCE. *Heat Treating and Forging* 16, 1150-8 (1930).—The av. processing of an extra-deep drawing auto-body sheet as made in this country and in Germany is compared.

J. BALOZIAN

Characteristics of rimmed steel. EDWARD C. BITZER. *Blast Furnace Steel Plant* 19, 249-51 415-7 422 (1931).—Rimmed steel is one which has been partially deoxidized either in the furnace or in the ladle and is poured in this condition. C, Mn and S have a profound effect upon the rimming properties of the steel. Too much Mn and C will cause the steel to grow without any indication of rimming action. 0.06% S would cause the rimming action in an ordinary open hearth heat to be too weak and the blowholes in the ingots would be too near the surface. Al facilitates the pouring of certain grades of rimmed steel whose C is low. It is added in shot form according to the judgment of the pourer. It makes the metal pour quietly, thus cutting down the pouring time. Segregation is the principal evil in rimmed steels. Since different heats of rimmed steel segregate to different degrees, it seems logical that a more detailed knowledge of the effect of FeO content in the metal would lead to the mfg. of a better grade. A method has been developed for detg. FeO in liquid steel. The importance of temp. in the mfg. of rimmed steel cannot be over-emphasized, as excessive temp. has probably ruined more steel than the total amt. spoiled from other causes. The advantages of rimmed steel are its low cost and use of less deoxidizer, it contains a min. of non-metallic inclusions, the yield from ingot to finished product is greater because of the absence of pipe in properly made steel. Better control is needed to decrease segregation, thin skinned ingots and over oxidation.

H. C. PARISH

Crystal structure of martensite. E. ÖHMANN. *Nature* 127, 270-2 (1931).—The present knowledge of tetragonal martensite is summarized and results of earlier investigators are reviewed with reference to C content and lattice dimensions. Ö employed Cr K radiation and obtained photographs of quenched steel specimens in which the line (101) of the tetragonal phase is sep'd from the γ -Fe line (111). It was thus possible to det. the lattice dimensions of the tetragonal structure. Results are in agreement with those of Kurdumoff and Kaminsky that the tetragonal martensite is a supersat'd soln of C in α -Fe. Ferrite and tetragonal martensite are thus one and the same phase, but as they are often present in one and the same specimen as sep. constituents, it seems convenient to denote the tetragonal phase as α' . Westgren and

Phragmen suggested that the C atoms dissolved in γ -Fe do not occupy any points of the face-centered lattice, but are statically distributed in the interstices between the Fe atoms. Seljakow, Kurdumoff and Goodtzow suggested that the same might be the case in the α' phase, but O thinks this improbable as the space available for the C atoms would be extremely small. Another probable explanation is that there is a complex substitution in such a way that a group of 2 C atoms is substituted for 1 Fe atom in the lattice. These compds. have a tetragonal structure in which the C_2 group is oriented parallel to the tetragonal axis, an explanation which holds equally well for tetragonal martensite. By density measurements, it is shown that the assumption of a complex substitution of C atoms is in good agreement with x ray intensities. The only suggested structure of tetragonal martensite which explains the observed density, the increase of vol. with the C content, and the elongation of one of the crystallographic axes may be described as follows. In the body-centered lattice groups of 2 C atoms statically distributed replace some of the Fe atoms. The C atoms are most probably oriented in such a way that the axes of the C_2 groups are parallel to the tetragonal axis of the lattice.

G. L. CRAIG

Effects of alloys on rolled and cast steels. R. C. GOON. *Iron & Steel Can.* 13, 121-2, 141-2 (1930). *Can. Machinery* 41, No. 16, 39-40.—A brief summary is given of the reasons for using Cr, Ni, Co, Mo, Mn, V and Si, individually and in certain combinations as alloys, in rolled and cast steels.

DOWNS SCHLAAP

Some lesser known facts concerning alloy steels. J. H. ANDREW. *Trans. Inst. Eng. Shipbuilders Scotland* 73, 166-98 (1929).—See C. A. 24, 810.

A. N. H.

Relative merits of some different alloy steels with respect to certain mechanical properties. BRADLEY STOUTON and WILBER E. HARVEY. *Proc. Am. Soc. Testing Materials* 30, Pt. II, 241-57 (1930).—Data in the literature are studied by the application of various "quality factors," such as tensile times elongation, elastic limit times reduction of area, tensile times Izod, etc., in an effort to find a quality factor that will pick out superior steels. The data are tabulated in various ways. In one classification such widely varying compns. as C 0.47, Cr 0.51 and C 0.09, Cr 16.0% are classed together as Cr steels. Variations in methods of testing for elastic limit, in notches used on Izod bars, etc., in size of bar from which a test specimen was taken, whether heat treated in test bar or larger size, etc., are not given and these variations probably affect the tabulations. No attention is paid in quality factors employed, to cleanliness of steel, ease of rolling, response to heat treatment, depth hardening, weldability, machinability or cost, though attention is called to the fact that some of these are important. Numerous data are recorded on Ni, Ni-Cr and Cr-V steels of tensile strengths from 100,000 to 300,000 lb./sq. in. with references to the original sources of the data. No conclusions are drawn as to which, if any, "quality factor" is to be preferred.

H. W. G.

Can residual analysis by decomposition with chlorine be applied to alloy steels? ROLAND WASMUTH. *Arch. Eisenhüttenw.* 4, 155-9 (1930).—In order to apply this method to the detn. of metal oxides the alloy and any metal carbides which may be present must be completely decomposed at temps. at which the oxide is not yet attacked by the Cl. A no. of alloys and metal carbides and also a number of oxides, by themselves as well as in the presence of C, were investigated for their resistance against Cl at various temps. The oxides of W, Mo, Ni, Co cannot be detd. by this method, as they are attacked by the Cl at temps. at which the carbide or the metallic alloy are not yet completely decomposed. It seems possible to det. Cr oxide if a decompn. temp. of 500° is chosen. Such residual analyses were made on Cr steels and showed that the materials contained varying quantities of Cr oxide, from which conclusions could be drawn as to the behavior of the alloy. The use of chips or metal pieces as initial material for the analysis is briefly discussed.

LEOPOLD PESSEL

Alloy steels in the railroad field. CHARLES MCKNIGHT. *Trans. Am. Soc. Mech. Eng., Railroads* 52, 81-93 (1930).—There is an increasing interest in the use of alloy steels for railroads. They are used purely as a matter of economy. Better work must be done at the same cost or the same work at less cost. The use of alloy miscellaneous castings materially cuts down weight with the same strength and same cost as C steel castings. C steels are used for special applications. Semi-alloy steels are slightly more expensive than C steels and have slightly better properties. Some of the uses include boiler materials, forgings and castings. Corrosion-resisting materials, nitriding and miscellaneous materials are discussed and the economies of each are given.

R. RIMBACH

Preliminary report on the behavior of copper steel rails. FRANCESCO ABBOLITO. *Rivista tec. ferrovie ital.* 37, 292-6 (1930).—Cu steel rails (contg. 0.08-0.23% Cu and having 66-83.5 kg./sq. mm. strength) were examd. after 16 yrs. service: a loss in the

breaking strength of 14 kg was found and an av. loss of 4.2 sq cm on the original section of 47 sq cm. G. A. BSAVO

Results of laboratory tests and practical experience with rails on Swiss railways. P. LUCCHINI AND M. ROß. *1st Communications New Intern Assoc for Testing Materials* 1930A, 190 212 (in German).—Simultaneous tests in the Swiss Federal Testing Laboratory and the Swiss Federal Railways on various mild steel rails of known compn., especially with regard to wear, showed that hard steel rails and heat treated rails of great hardness have a greater resistance to wear than older and newer rails of lower hardness and strength. F. M. SYMMES

Alloy cast iron and high-strength cast iron in the United States of America. F. B. COYLE. *1st Communications New Intern Assoc for Testing Materials* 1930A, 35-48, of C. A. 23, 4175.—The extent of development of alloy cast Fe in the U. S. A. is described. The alloying elements are carbide forming and graphitizing. There is described the general effect of alloying elements on grains, tendency to chill, hardness, machinability, resistance to wear, d. and strength, and the strength limits of plain and ordinary alloy Fe are given. To develop extremely high strength, superheating, partial malleabilization, low C cupola Fe and white cupola Fe graphitized by ladle adds may be used. F. M. SYMMES

Alloy additions improve gray iron pressure castings. F. J. McGRATH. *Foundry* 59, 52-5 (1931).—Ni and Cr alloys in the form of shots are sprinkled into molten metal. These 2 metals eliminate the use of expensive chills, and the grain structure is uniform to the center of the test pieces. Photomicrographs are shown of test pieces which broke around 5000 lb. on the transverse test. H. C. DUES

The relations between structure and properties in metallic alloys. A. GUESTLER. *1st Communications New Intern Assoc for Testing of Materials* 1930A, 320 7.—A philosophical discussion of individual crystals of pure metals, of compds., of solid solns., and of aggregates. H. W. GILLET

Heat-resisting alloys in furnaces. L. J. STANBURY. *Heat Treating and Forging* 16, 1031-5, 1191 2 1195 (1930).—The principles involved, the effect on the properties of the alloy of various chem. elements and of heat treatment are discussed. J. B.

The influence of silicon and manganese on the type of solidification of iron-carbon alloys. OTTMAR VON KEIL AND FRANZ KOTIZA. *Arch Eisenhüttenw.* 4, 295-7 (1930).—The influence of Si and Mn on the type of solidification and the carbide decompn. of Fe-C alloys is investigated metallographically and by cooling-curve detns. The fusions were made at 1350-1400° and the cooling velocity was 30°/min (av.). At these cooling velocities Fe-C-Si alloys of hypo-eutectic compn. show metastable solidification with subsequent carbide decompn., and only in the hypereutectic alloys does the stable solidification appear. By addn. of Mn the white solidification is more stable at low C contents, while at approx. eutectic concns. the stable phase is favored. Also, the field of metastable solidification with subsequent decompn. is smaller with increasing Mn content. J. BALOZIAN

The alloys of iron and manganese. V. N. KRIVONOX. *Heat Treating and Forging* 16, 1538-40 (1930).—The heat treatment and metallography of Fe-Mn alloys are discussed. Micrographs (9) are shown. J. BALOZIAN

Dimensional stability of heat-treated aluminum alloys. J. D. GROGAN AND D. CLAYTON. *J. Inst. Metals*, Advance copy No. 553, 28 pp., *Engineering* 131, 371-3 (1931).—It was found that the change in length of both cast and forged heat treated Al alloys over a period of years was inappreciable. Changes in length on machining these alloys however, may occur. Samples of Y alloy, duralumin and 25S 3 in. in diam. by 5 in. long were heat treated and machined into dumb-bell form. As the outer surface was removed appreciable shrinkage occurred in the fully aged alloys. The shrinkage was greatest in those quenched in cold water, less in those quenched in oil and almost negligible in those quenched in boiling water. The decrease in length in some cases amounted to almost 0.020 in. Tensile and hardness tests indicated that the alloys quenched in boiling water were as effectively hardened as those quenched in cold water. Dimensional changes on machining did not occur in annealed alloys. The corrosion resistance of the hot water quenched alloys may be low, and the greater resistance of the cold-water quenched alloys may result from compressive stresses in layers near the surface, which heal intercryst. corrosion cracks. J. L. GREGG

A note on the silver-rich aluminum-silver alloys above 600°. T. P. HOAR AND R. K. ROWNTREE. *J. Inst. Metals*, Advance copy No. 555, 6 pp. (1931).—Alloys up to 15% Al were studied by thermal and microscopic methods. The results substantiate Petrenko's diagram except that the diagram is shifted to higher temps. The diagram is similar to the Cu-Al diagram. J. L. GREGG

Influence of heat treatment on resistance of aluminum-silicon alloys containing up to 2.5 percent silicon. L. GUILLET AND M. BALLAY *Rev metal* 27, 398-403 (1930) —See C. A. 24, 5702 E. J. C.

An air-hardening copper-cobalt alloy. CYRIL S. SMITH. *Mining and Met* 11, 213-5 (1930) —A Cu alloy containing between 1 and 5% Co when quenched sufficiently rapidly from a high temp remained in the soft alpha condition. The quenched alloy on re-annealing at a suitable temp was hardened by submicroscopic pptn of the Co-rich solid soln. in a manner similar to duralumin. The alloy softened again if the annealing temp was too high, as a result of the coagulation of the ppt. A peculiar scale resulted when the alloy was annealed in an oxidizing atm. Underneath the usual black and red scale, there appeared a layer of tightly adherent metallic "sub-scale" which was detached only when the metal was very severely distorted. It consists of polygonal grains continuous with the underlying metal but containing innumerable black particles of Co oxide. F. W. BORECH

Rolling of alloys of copper and phosphorus containing up to 5 percent of phosphorus. OWEN W. ELLIS. *J Inst Metals*, Advance copy No 551, 6 pp (1931) —Contrary to recent reports, Cu contg as much as 5% P can be rolled. Sheets for brazing have been rolled to 0.015 in. The Cu P alloys were rolled at a temp of 450° or above to between 0.02 and 0.04 in. and then rolled cold. A strip of the 5% P alloy 0.018 in. thick had a tensile strength of 88,000 lb/sq in. and practically no elongation. J. L. GREGG

Nickel alloys. W. R. BARCLAY. *Metal Ind* 37, 414-6, 437-40 (1930) —A historical review of B's 35 years of personal experiences in the manu of Ni alloys. The importance of const. study and research is insufficiently realized. O. M. SMITH

Etching nickel and its alloys. W. A. MUDGE. *Metal Progress* 18, No 6, 72-3 (1930) W. A. MUDGE

The advantages of lead-antimony and lead-antimony-cadmium alloys over lead in the manufacture of sheaths for telephone cables. H. GODIN. *Rev univ mines* [8], 5, 106 (1931) —The failure of cable sheath has been found to be largely a matter of fatigue. On the basis of this property, an alloy of Pb 99.25, Sb 0.5 and Cd 0.25% has been found superior to the 1% Sb alloy and very much better than Pb. R. S. DEAN

Special bronzes. WILLY CLAUS. *Metallborse* 20, 1769-70, 1825-6, 1881-2, 1937-8, 1993-4, 2049-50, 2105-6, 2161-2, 2217-8, 2274-6, 2331-2, 2386-7 (1930), cf C. A. 23, 5144. —Special bronzes are defined as alloys contg at least 78% of Cu and other metals besides Zn. The special bronzes discussed are those contg Pb, Mg, Cd, Si, Mn, P, Al, Ni, Sb, As, Bi, Co, Ag, Fe, Cr, W, V, Mn, Ti and Be. A complete discussion of the properties of these various special bronzes is given but no new data. R. S. DEAN

Carburizing with gas in underground retorts. W. F. McNALLY. *Metal Progress* 19, No. 1, 39-43 (1931) —Gas carburization with natural gas, propane and dehydrated city gas is superior to solid material carburization because of lower labor, fuel, carburizing media and container costs together with greater uniformity and flexibility. W. A. MUDGE

Pickling of metals. E. GIRARD. *Rev chim. ind* 40, 7-12 (1931). —A review P. THOMASSET

Some practical features of wire pickling. W. G. IMHOFF. *Wire* 5, No 3, 77, 79-80, 94-5 (1930). E. I. S.

How to select thin plate for deep drawing. H. T. MORTON. *Metal Progress* 18, No. 6, 54-9 (1930) —A discussion is given of chem. compn, slag inclusions, surface decarburization and annealing upon drawing quality. W. A. MUDGE

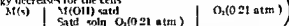
Coatings, corrosion protection and coloring of aluminum. HERBERT KURREIN. *Metallwirtschaft* 9, 1009, 1031-2, 1049-50 (1930) —A review LEOPOLD PESSER

Theory of the rust-protective action of protective coatings. W. BECK. *Korrosion u. Metallschutz* 6, 229-30 (1930), *Gas J* 193, 216-7 (1931) —By means of colloid-chem. considerations an attempt is made to explain the mechanism of the protective action of protective coatings. An explanation is given of why, in spite of a very distinct lack of elec. insulation, an effective rust protection may exist. Adsorption layers may be formed on the surface of the metal which may have great protective value and yet, on account of their great thinness, only a very small elec. resistance. The impossibility is emphasized of drawing conclusions from the poor insulating qualities of the protective coating as to its protective value. LEOPOLD PESSER

Selective solution of magnesium corrosion products. L. WHITBY. *J. Soc. Chem Ind* 50, 83-5T (1931) —A boiling 20% soln. of pure CrO_3 may be used to remove corrosion products rapidly from Mg and Mg-base alloys. The soly. of the metal in acid is negligible for most purposes. If Cl^- is present its activating effect may be inhibited.

by adding small quantities of Ag_2CrO_4 to the CrO_3 soln. SO_4^{--} exerts no activating action when present in small quantities. B. E. ROETHLI

Initial corrosion rates of metals. R. H. BROWN, B. E. ROETHLI AND H. O. FORREST. *Ind. Eng. Chem.* 23, 350-2 (1931).—The initial corrosion rates of Al, Zn, Fe, Cu, Ni, Sn and Ag in oxygenated water were studied. All the metals investigated were shown to have finite corrosion rates. Tendencies of the metals to corrode (calcd. as free energy decreases for the cells



were obtained and found to be in better agreement with the facts than tendencies obtained from consideration of the mofal e m f series. Corrosion rates decreased with immersion time. This was attributed to the formation of films which reduced the rate of diffusion of ions and dissolved O_2 . B. E. ROETHLI

Metal corrosion. G. CHAUDRON. *La Nature* 1930, 1, 172-4; *Wasser u. Abwasser* 27, 171.—The electrochem. corrosion theory as applied to metals and the importance of O_2 in corrosion are explained. Corrosion is best detd. on metals by detg. the loss of wt. during exposure under fixed conditions of temp., etc. Protective alloys and coatings are discussed. C. R. ILLPAS

Influence of stress on corrosion. D. J. MCLOAM, JR. *Am. Inst. Mining Met. Eng. Tech. Pub. No. 417*, 39 pp (1931). cf. *C. A.* 24, 2412.—Special attention is given to Ni, Al bronze, stainless Fe, nitrided steel and Muntz metal. Each expt. was in 2 stages: (1) a corrosion stage, in which the specimen was corroded with or without cyclic stress; (2) a fatigue stage, in which the specimen was tested to fatigue failure and the fatigue limit was estd. The lowering of the fatigue limit is used as a measure of the "damage" due to corrosion. Diagrams of 3 types are presented: type 5 illustrating relative influence of stressless corrosion and corrosion under cyclic stresses in causing damage; type 10 illustrating influence of cycle frequency on net-damage; and type 11 illustrating influence of stress range. These diagrams, in addn. to diagrams for steels, Al alloys and Monel metals presented in previous papers, illustrate the behavior of a great variety of metals under similar conditions of corrosion. The rate of net damage varies as the third to at least the fifth power of the corrosion stress, depending on the metal, cycle frequency and corrosion conditions. The general conditions favoring intercryst. corrosion are discussed. Application of stress-corrosion data to design, construction and operation of machinery and structures is discussed briefly. C. L. MANTILL

The effect of artificial aging upon the resistance of superduralumin to corrosion by sea water. K. L. MEISSNER. *J. Inst. Metals*, Advance copy No. 557, 15 pp; *Engineering* 131, 400-8 (1931). cf. *C. A.* 25, 62.—Specimens of superduralumin (0.8% Si) that had been quenched from 500° and aged for 20 or 40 hrs. at temps. up to 200° were exposed in the North Sea for 3 months. During exposure they were under water for 7 or 8 hrs. and in the air for 4 to 5-hr. periods. The corrosion was detd. by appearance, change in weight, change in tensile properties and microscopic examn. The samples aged at 125° and 145° were badly pock marked. Tensile tests indicated that these samples had been most severely attacked. Microscopic examn. showed that these samples were corroding at the grain boundaries and that the pock marks were caused by chunks of Al falling off. In such an alloy there is therefore a crit. tempering range between 100° and 150° at which the corrosion resistance is at a min. This range is just below the crit. hardening temp. J. L. GREGG

First report of the corrosion of pipes sub-committee. ALBERT STOKES, et al. *Inst. Gas Eng.* 1930, *Communication No. 20*, 10 pp.—The effects of various coatings on wrought Fe and steel were detd. in 2 types of tests: (a) The coated metal was placed in soil and made the anode of a closed cec. system. The source of current was a 6-v. battery. The current was measured at different intervals of time and the resistance of the coating calcd. (b) The e m f differences between a coated and uncoated specimen were measured at various times. The resistance of the coating decreased with increasing H_2O content while the e m f measurements showed increasing values. Painted pipes showed much greater resistance to the flow of current than bare pipe. Pipes covered with paper wrappings and paints were about 50 times as resistant to the flow of current as paint coatings. Deep pits occurred in the metal surfaces where the coatings were broken. Cement coatings were found temporarily resistant to stray current electrolysis. Also in *Gas J.* 192, 725-9 (1930). B. E. ROETHLI

Error of averages and its application to corrosion tests. C. H. HUMES, R. F. PASSANO AND ANSON HAYES. *Proc. Am. Soc. Testing Materials* 30, Pt. II, 448-55 (1930).—If all samples of a given material in a corrosion test failed at the same time,

only one sample would be required. When failures occur over a time range the no and area of samples that must be tested to give a true measure have to be detd by the methods of statistics. The mathematics of the procedure needed to arrive at the no of samples required for reasonable assurance of a given desired accuracy in the mean of the exptl data are discussed. H W GILLET

Outline of welding and allied processes. F T LLEWELLYN. *Metal Progress* 18, No 6, 95-104. 19, No 1, 51-60 (1931). —This is an outline of the practice recommended by the Am Soc Steel Treating comm. W A MUDGE

Electric arc welding of aluminum. W M DUNLAP. *Welding* 1, 762-6 (1930); cf *C A* 24, 2975-5277. —A discussion of equipment, flux, electrodes, procedure control and advantageous application. E I S

Metallurgy of the fusion welding process. F RAPATZ. *Stahl u Eisen* 51, 245-53 (1931). —The fusion welding process is reviewed. The desirable qualities of welding rods are: good melting off, that is the rod should form uniform drops in the arc and the cast should be free of pipes, then good fusing together, that is a good alloying of the weld with the rest of the material, the deeper this "burning in" the better the weld. For elec welding if d-c is used with the electrode as the neg pole, bare rods or wires are required contg non-metallic admixts best in the form of non-metallic cores. This method is best suited for soft ingot steel. For steel with higher C content and alloy steels this welding on the neg pole cannot be used. If the pos pole welding is used for all types of alloyed steel, good melting off of the rods is obtained. However, the alloying, the "burning in," is poor. This method is best suited for high C steel. For a-c welding rods with an outside cover, consisting of either asbestos or a mixt. of Fe oxide, materials giving an exothermic reaction and metals readily combining with O, such as Al or ferromanganese, are used. Wires or rods with non-metallic core can be used also. Plain steel rods or wires cannot be used. In d-c welding, with well-covered wires or rods, the electrode should be the pos pole. In gas fusion welding the rods should be free of non-metallic admixts. The compn of the weld depends on the compn of the welding rods, on the gas absorption and oxidation of the various constituents of the material welded and on the alloyed "burnt in" layer. The effect of the alloying elements, such as Ni and Cr on the characteristics of the welded part is discussed. There is an overheated zone around the weld, then a normal annealed, a slightly affected and in some cases, a blue brittle zone. The area of this heat affected zone is greater in gas fusion welding than in elec welding. The Brinell hardness figures for the various zones are given for both cases. In highly alloyed steel a very decided effect on the area around the weld is obtained, because of air hardening. On account of the zones of different hardness, great care should be taken in the evaluation of the results of the various bending and tensile tests. J A SZILARD

Valves for blast furnace and producer gas (JOHANSSON) 21. Corrosion reduced by new constructional materials (LEE) 23. Boiler tube corrosion halted by hot-process treatment (COUGHLIN) 14. The coking industry and its development in relation to the manufacture of Fe and steel (DICKIE) 21. Oxidation of pig Fe during its transformation into steel (MIKOHYAMA) 4. The critical anisotropic point of ferromagnetic crystals (Akulov) 2. Special refractories for metallurgical research (TURNER) 19. Piping materials for mineral water (MAURER) 14. Special loading of long submarine cables (Brit. pat. 339,049) 4. Enamelled boilers and other vessels with heating or cooling coils in their walls (U. S. pat. 1,796,123) 1. Attaching rubber to metals (Brit. pat. 339,421) 30. Activating metals or alloys (Austrian pat. 121,243) 3. Combined gas-producer furnace for heating metals (Brit. pat. 338,893) 21. Pb pigments (Brit. pat. 339,357) 26. Charcoal for Fe metallurgy (Hung. pat. 101,960) 21.

KEHL, ROBERT J. *Oxy-acetylene Welding Practice*. Chicago: Am. Technical Soc. 104 pp.

KROEL, H. AND NIESE, H. *Die elektrischen Schweissverfahren*. Berlin: Walter de Gruyter. 135 pp. M. 1.80. Reviewed in *Chimie & Industrie* 25, 533 (1931).

ROUX, A. *Agenda Dunod. Metallurgie, 1931*. Paris: Dunod. 496 pp. F. 20.

OSANN, BERNHARD. *Lehrbuch der Eisen- und Stahlgiesserei*. 5th ed., revised. Leipzig: Wilhelm Engelmann. 693 pp. M. 21; linen, M. 24.

Ore flotation. A. BORSIG G. M. B. H. *Ger.* 515,757, July 24, 1926. Addn. to 446,086. App for sweeping out the slime from clarifying vessels is described.

Flotation apparatus for concentrating ores, etc. FRANK F TLMORF Ger. 518,301, Aug 13, 1927 See Brit 275,778 (C A 22, 2296)

Flotation apparatus for concentrating ores. LÉON C V BENETT Fr. 37,243, July 22, 1929 Addn to 563,182

Extracting metals. RUDOLF SCHRAUK and KLEMENS SCHUMACHER. Austrian 121,535, Oct. 15 1930 In extg metals from oxide ores by blowing reducing gases through the fused ore, e g., in effecting the reaction $\text{FeO} + \text{CO} = \text{Fe} + \text{CO}_2$, the metal is continuously withdrawn from the reaction, in order to shift the equil point in the desired direction This may be effected by maintaining such a reaction temp that the ore remains liquid but the metal solidifies Also a converter may be used to which the reducing gases are fed at points above the bottom, the metal sinking to the bottom out of contact with the gases Powd C may be blown in with the gases

Extracting metals RUDOLF SCHRAUK and KLEMENS SCHUMACHER. Austrian 121,523, Oct 15, 1930 In the extn of Fe and other metals by blowing highly heated reducing gases through molten ore, the temp of the ore is maintained by passing an elec current and (or) high frequency current through it Suitably, the ore is melted in a hearth furnace and then transferred to a tiltable elec furnace to which the reducing gases are supplied

Treating ores or concentrates of copper and nickel, etc. E A ASHCROFT. Brit. 338,556, Aug 20 1929 Material contg Cu or Ni sulfides or both, or Cu or Ni oxidized compds together with added S such as Fe sulfide, is treated with air or O at 600° to convert the Cu and Ni to sulfate (without disturbing or rabbling the materials during the reaction) The heat of reaction is usually sufficient to maintain the temp after the reaction is started, and the charge may be covered by a thin layer of inert material When Ni is present FeSO_4 or FeCl_2 may be added with advantage Most of the Fe present remains in the residue as oxide on leaching, together with sulfates of gang materials such as CaSO_4 O or air enriched with O may be used to start or assist the reaction, and various other details and modifications of app and procedure are described

Leaching copper and nickel sulfide ores I BALINT Hung 101,408, Feb 15, 1930 The ores are leached at a temp higher than 100°, under pressure, with FeCl_3 or CuCl_2 as solvent

Leaching copper sulfide ores. I BALINT Hung 101,405, Oct. 5, 1929. The used CuCl soln may be regenerated by sepg it into part (1), contg as much Cu as was dissolved out of the ores, and part (2), contg as much as was originally used for leaching Cu is sepd from part (1) by cementation The remaining soln is united with part (2) and the mixt transformed to CuCl_2 by aeration The iron hydroxide ppt formed in the cementation can be sepd

Combined reduction of ores and hydrogenation of carbonaceous materials. J. L. FORLEN Brit. 339,276 July 25, 1929 Ores such as Ni ore are reduced to metals, and carbonaceous materials (such as lignite and its distn products) are hydrogenated to form volatile fuels, by first reducing the ore in the presence of the carbonaceous material or its distn products and H heated to about 350–400°, then allowing the temp to fall to 150–300° and hydrogenating the gaseous products of the reaction by contact with freshly reduced metal An arrangement of app and various details of procedure are described, and the process may be applied to the production of hydrogenated products from anthracene etc

Treating sulfide ores NATIONAL PROCESSES, LTD., and T. B GYLES Brit 338,886 Aug 27, 1929 Sulfide ores are blast roasted or sintered together with a coarse granular material, such as sintered ore, and a small porportion of sulfate or sulfite of a metal other than Fe or an alkali or alk earth metal (suitably sinter from a previous operation treated so that it contains ZnSO_4 with or without ZnSO_3) Sulfurous gases may be passed over the sinter (preferably moistened) as described in Brit 321,390 (C A 24, 2515) or the gas may be passed over the mixt of raw ore and diluent in a mixer from which it is fed to the sintering app (or ZnSO_4 liquors may be used instead of sulfurous gases)

Direct production of metals such as iron from ores. THOMAS ROWLANDS. U. S. 1,795,681, March 10 Ore such as Fe ore is passed through preheating, reduction and cooling zones of a rotary furnace and carbonaceous fuel such as coal is subjected to low temp carbonization in the cooling zone by heat from the treated charge that has passed through the reduction zone but without direct contact with the charge; ore reduction in the reduction zone is effected at comparatively low temp. and considerably below melting temp by the reducing gases and solid fuel from the low temp carbonization App is described

Production of pig iron from ores poor in iron and from iron-containing bauxites.

f PFEIFER Hung 100,509, March 12, 1929. Ores contg 30-35% or less Fe are mixed with slag forming materials, *e g.* silica or alumina, and heated in the presence of CO or hydrocarbons. The iron produced is satd with C and the m. p. is thus lowered to 1150-1200°. The slag forming material should be so selected that the slag also melts at this temp and thus can be easily sepd from the melted iron. For Fe-contg bauxites, the slag forming material should contain alkalis, *e g.* Na_2CO_3 , which lowers the m. p. of alumina by formation of Na aluminates.

Direct production of cast iron from ore. **J E FLETCHER** Brit 338,535, Aug 19, 1929. Reduction is effected in one or more primary rotary furnaces and the molten product is passed into a secondary furnace in which it is further heated with any accompanying slag and with desired addns of modifying or alloying materials. Some mention is made of the applicability of the process to production of metals other than iron.

Preparing cast iron from ores. **A G KHOTYAN** Russ. appl 24 001, Mar 27, 1928. A mechanical device is specified.

Extraction of iron from its ores. **VEREINIGT STAHLWERKE A G** Ger 518,392, Sept 17, 1927. See Fr 35,306 (*C A* 24, 3746).

Reducing oxide ores of metals such as iron. **SAMUEL L MADORSKY** (to Gathmays Research Corp.) U S 1,796,871, March 17. A reducing gas is passed through molten ore in part of a furnace (details of construction of which are described), and the waste gas passing from the ore is passed over molten ore in another part of the furnace to effect partial reduction of the ore. Cf *C A* 24, 4499.

Titanium recovery from ores. **OSCAR T COPPELT** U S 1,795,361, March 10. A Ti ore contg Fe is digested with acid such as 75-90% H_2SO_4 and the sol. salts thus formed are dissolved, the Fe and some of the Ti are reduced (suitably by use of Fe or Zn) and an NH_4 halide such as NH_4Cl is added in an amt. equal to the Ti present, water is added to double the volume, followed by boiling and maintaining the diln until a ppt. contg Ti in combination with O is formed, and this ppt. is removed.

Tungsten from ores. **CHARLES V. IREDELL** (to Westinghouse Lamp Co.) U S 1,796,026, March 10. Alk. solns such as may be formed by treating a W ore with alkali are treated to reduce their alkalinity to 0.3-0.7% free alkali (using phenolphthalein as indicator), and a Ca salt such as CaCl_2 is then added to the soln. to ppt. the W content. A product of the formula $\text{CaWO}_4 \cdot \text{H}_2\text{O}$ may be obtained.

Acid extraction of metals such as vanadium and zinc from ores. **ROYALE H STEVENS, GERALD C NORRIS and WM N WATSON** (to Rhodesia Broken Hill Development Co., Ltd.) U S 1,795,412, March 10. The ore is leached with a soln. of relatively low acidity such as a spent electrolyte and the residue is sepd. and re-leached with acid maintained at a concn. equiv. to not less than 4% H_2SO_4 .

Reducing zinc, lead and like ores. **FREDERICK L DUFFIELD** U S 1,796,562, March 17. See Brit. 310,252 (*C A* 24, 531). U S 1,796,563 relates to an app. similar to the one described in the preceding patent.

Apparatus for reducing ores such as those of zinc. **HENRY E COLEY** U S 1,797,130, March 17. Ore is fed into an internally heated reducing chamber provided with an outlet for gang and an inlet for introduction of hydrocarbon material to the ore in the reducing chamber, arranged so that the hydrocarbon material reaches the ore in substantially undecomposed condition, an inlet for combustion air is provided to effect combustion in the reducing chamber of products produced by the cracking of the hydrocarbon material. Various features of construction are described. Cf *C A* 24, 2711.

Obtaining zirconia from ores. **CHARLES O. TERWILLIGER** (to Ward Leonard Elec Co.) U S 1,796,170, March 10. Finely ground ore is treated with H_2SO_4 and HCl of such strength as to dissolve impurities without dissolving the zirconia, and the soln. is removed and the residue is washed.

Treating metals in ladles (as in an induction furnace). **F. KRUPP A-G** Brit 339,579, April 30, 1929. When heating molten steel or other metal or alloy in a ladle by the heater described in Brit. 312,063 (*C A* 24, 760), the ladle is placed in a chamber in which a vacuum is created. Cf *C A* 24, 4468, 4502.

Mold for metal casting. **MASCHINENFABRIK ADOLF VIEBACH** Ger. 502,935, Feb 3, 1929.

Mold for making tubes by centrifugal casting. **SOCIETÀ ALTI FORNI, FONDERIE, ACCIAIERIE E FERRIERE FRANCHI-GREGORINI** Austrian 120,888, Aug 15, 1930.

Vertical mold for making hollow bodies by centrifugal casting. **HEINRICH LENT, HERMANN J. SCHIFFLER and GUSTAV TICHY** Ger. 518,122, Mar. 26, 1929.

Core for centrifugal casting molds. SOC. ANON. DES HAUTS FOURNEAUX ET FONDERIES DE PONT A MOUSSON. Ger 518,123, Nov 6, 1926

Machines for making hollow bodies by centrifugal casting. SOCIETÀ ALTI FONDI, FONDERIE ACCIAIERIE E FERRITAT FRANCHI-GALCOINI. Austrian 120,880, and 120,880 Aug 15 1930

Two-part metal founding molds. HENRY T. RICHARDSON. Ger 515,772, May 19, 1928

Cooling system for casting troughs such as those used in centrifugal casting. FRANK G. CARRINGTON (to Lurie Engineering Co.) U. S. 1,796,644, March 17. Various details of app. are described. U. S. 1,796,645 also relates to centrifugal casting app.

Water-cooled die-casting machine. COMPAGNIE GÉNÉRALE DES CONDUITES D'EAU. Soc. ANON. Ger 515,799, April 21 1929

Metal-casting apparatus. CARL F. BECKT. U. S. 1,797,041, March 17. Structural features

Centrifugal casting apparatus suitable for casting belled pipes. FRANK G. CARRINGTON (to Lurie Engineering Co.) U. S. 1,794,637 March 10. Structural features

Centrifugal casting of metal pipes. SOC. ANON. DES HAUTS FOURNEAUX ET FONDERIES DE PONT A MOUSSON. Brit 379,095, Dec 10, 1928. App. and various mech. details are described

Casting metal bushings. OTTO BEAGNER. Ger 515,801 Sept 20, 1928

Casting flanged pipes. HEINRICH NOLLE. Ger 515,830, Nov 10, 1928

Water-cooled slide-way for ingot-reheating furnaces. FRIEDRICH C. SIEMENS. U. S. 1,795,410, March 10

Bessemer converter. L. I. KONATOVSKII. Russ. appl. 67,344, Mar 30, 1930. Mechanical features

Apparatus for charging the Martin furnace. N. G. GRACFSEL. Russ. appl. 68,282 June 7 1930. Mechanical details.

Coal reversible gear for Martin furnaces, etc. O. REINER. Brit 338,788, Jan 10, 1929. Structural features.

Regenerative oven, especially useful as a metallurgical smelting furnace. WALTER ALBERTS and PAUL ZIMMERMANN. Ger 515,835, May 22, 1929

Oil-fired cupola furnace. PETER MARX. Ger 515,796, Feb 2, 1928.

Slag collector for cupola furnace. FRIEDRICH SCHINKT. Ger 518,170, Jan 28, 1927

Smelting metals such as iron and iron alloys. CARL BRACKELSHAG. U. S. 1,797,125 March 17. In operating rotatable drum furnaces for melting metals by use of powd. coal as fuel the fuel is burned directly in the melting chamber of the furnace by the aid of combustion air of such vol. and pressure as to cause the fuel to be burned in suspension and the ashes of combustion to be blown through the outlet of the furnace. App. is described

Apparatus for charging retorts such as those for zinc production. WALTER F. NAYLOR (to Beardsley & Piper Co.) U. S. 1,796,802, March 17. Structural features

Modifying the properties of metals. E. G. HENDERAT. Brit. 339,511, Aug 16, 1929

Metals (suitably after hardening by working or otherwise) or other substances such as rock salt, sugar or quartz are subjected to the action of a coned, locally generated magnetic field "or like flux or vibration other than heat" (the magnetic field or the like being alternating or rotating)

Fusing refractory metals. SIEMENS & HALSKER A-G. Ger 518,499, Nov 2, 1926

Metals such as Ta, W, Th or their alloys are completely fused in strongly cooled crucibles made of materials of lower m.p. than the said metals, e.g., of quartz, Cu or Ag. The metal to be fused may be heated by high frequency elec. current, and the crucible, if electrically conductive, may be made in a no. of insulated segments.

Metal-reduction bomb suitable for use with calcium and calcium chloride. WM. F. KIERNAN (to Westinghouse Lamp Co.) U. S. 1,795,789, March 10. Various details of stopper construction and fastening are described. The main portion of the bomb may be formed of an alloy resistant to heat and to oxidation such as "nichrome" or "hardite"

Coating metals. OTTO SPRENGER. PATENTVERWERTUNG JIROTKA M. B. H. Fr. 37,112 June 19 1929. Addn. to 655,786 (C. A. 23, 3894). The deposition of MnO_2 on Al, Mg or their alloys is accelerated by adding bi- or tri-valent salts of Mn or ferrous or stannous salts to the baths contg. the permanganate and free acid. Cf. C. A. 24, 2997

Galvanizing pot. O. MIDDERMANN. Brit 339,384, Sept. 20, 1929. Walls of iron

galvanizing pots which come in contact with the molten Zn are freed from pores by melting their surface layers (suitably with a welding burner or elec arc)

Rolling thin metal strips or sheets from bars. L. E. BROEMEL. Brit. 338,771, Dec. 18, 1929 Mech features.

Pointed tubular metal articles such as hypodermic needles. S. J. EVERETT. Brit 338,592, Aug 22, 1929 In a process involving swaging, rolling or drawing a tubular billet on a core of metal capable of extreme elongation and retention of tensile strength, such as Ni Ag or brass (as described in Brit 210,156), the cored tube is subjected to heat treatment to anneal the core and temper the tube while on the core. Numerous details of procedure for the treatment of different specified metals are given

Blow-pipe for cutting metals by fusion with the use of oxygen. W LEBERL Brit 339,161, Jan 15, 1930

Pickling bath for metals. JAMES H GRAVELL and ALFRED DOUTY (Gravell to American Chemical Paint Co) U S 1,796,839, March 17. See Can 302,397 (C A 24, 4253)

Bright-annealing metals. SIEMENS-SCHUCKERTWERKE A-G Brit 339,423, Oct 22, 1928 Various details of app are described relating to a chamber in which the treated metal is cooled in inert gas the expansion and contraction of which is provided for by connection with a bell gas holder

Apparatus for annealing sheet-metal strips in coils or bundles. HERMAN C CAPPEL and OTIS C GROSS (to Greer Steel Co) U S 1,796,082, March 10 Structural features

Case-hardening composition. ERNEST C MORFETT (to American Cyanamid Co) U S 1,796,800, March 17 See Can 302,409 (C A 24, 4253)

Heat-treating wire and strip material. L D WHITEHEAD Brit 339,155, Jan 6, 1930 Wire or strip which is to be subsequently galvanized or otherwise surface-treated is annealed or cleaned (or both) by passing it through a horizontal metal tube arranged in a furnace and contg molten metal, salts or other suitable liquid material (the 2 ends of the tube being up turned to retain the contents) Various details of app are described

Heat treatment of iron. LIVE-BELT Co. Brit 338,924, May 31, 1929. Malleable cast iron is heated to a temp (suitably about 800°) sufficient to cause the formation of Fe carbide, cooled or quenched again heated to a temp (suitably about 650°) below the C-combining temp, and finally cooled or quenched (after heating to above the galvanizing temp if the product is to be galvanized by dipping). The first heating may be effected in the presence of a carburizing agent to form a hardened surface on the article. Brit 338,925 relates to metal similarly treated which is highly resistant to abrasion

Deoxidizing Iron. RUDOLF SCHAUR and KLEMENS SCHUMACHER Austrian 121,530, Oct. 15, 1930 In deoxidizing Fe by blowing reducing gases through the molten metal, the reducing gases are strongly preheated, and the temp of the metal is maintained by electrically heating, e g, by means of high frequency current.

Cementation of iron and steel. HERMANN FREUDENBERG (to Roessler & Hasslacher Chemical Co) U. S. 1,796,248, March 10 Hardening of ferrous articles is effected by immersing them in a fused salt bath such as soda and NaCl to which is added finely divided C (suitably in the form of charcoal) and an alkali metal hydroxide which serves to enhance the hardening effect

"Slushing grease" suitable for treating iron or steel to prevent rusting ARTHUR H. BRANSKY (to Standard Oil Co of Ind). U S 1,795,993, March 10 A wax such as montan wax is used with Na soaps of sulfonic acid compds derived from the treatment of hydrocarbon oils with strong H₂SO₄ and with lubricating oil

Apparatus for heat treatment of steel. V. A. MOZHAROV and I. V. MOZHAROV Russ appl. 20,409, Oct. 5, 1927; 32,405, Sept. 10, 1928 and 50,854, July 9, 1929

Tempering steel strip. ALFRED M REMINGTON (to Simonds Saw & Steel Co) U. S. 1,796,304, March 17. The stock is fed slowly through a hardening furnace and the forward end of the stock is fed rapidly into a drawing furnace a distance approx equal to the effective length of the furnace. App is described.

Special steel for car wheels, brake shoes, etc. H. J. VAN ROVEN. Brit. 338,528, Dec 8, 1928 Rails, wheel tires, etc., subjected to hard wear, are made of eutectoid steel contg double carbides with laminated pearlitic structure. The laminas may contain Fe carbide 75-85 and carbides of metals such as Mn, Cr, W, Mo and V 15-25%.

Steel billets. SAMUEL B. SHELTON. U S 1,795,519, March 10 An ingot is formed having an alloy skin which may contain Al alloy and the skin is subsequently removed so that surface imperfections are avoided in the billets formed from ingots thus prepd

Manganese steel TAYLOR-WHARTON IRON & STEEL CO Brit. 339,521, Dec 31, 1929 For refining the grain structure and improving the properties of Mn steel, it is heated for 12-46 hrs at 400-600°, slowly cooled, reheated to about 1000° and quenched. Various composit. of steels suitable for this treatment are described.

Heat treatment of alloy irons and steels P. R. KUPFENBACH and DARWINS, LTD. Brit. 338,912 Aug 28, 1929 Articles of alloy iron and steels contg Cr 8-20 and C 1.0-3.5% with or without up to 6% Co and up to 1% each Ni, V, Mo, Si and Mn and contg less than 0.05% each of S and P are hardened and rendered "stainless" and rustless by heating to 1150-1250° and then cooling at a definite rate corresponding to the rate at which steel in sheet form of about 18 gage cools in water at normal temp. Various details are given for treating spoons or other articles.

Alloys SIEMENS & HALSKER A. G. (Bruno Letkenheuer, inventor) Ger. 515,798 May 16, 1929 A speculum for medical use is made of an alloy contg Co or Ni 20-60 W or Mo 15-50 Cr 30 and C Si or B 1-5 parts.

Alloy for armoring alternating current cables. FELTEN & GUILLEAUME CARLSWEDE A. G. Brit. 339,383 Sept 25, 1929 See Ger. 504,318 (C. A. 24, 5281).

Alloys for sheathing electric cables. VEREINIGTE TELEPHON- UND TELEGRAPHIEN-FABRIKEN A. G. CZEHA NISSE & Co. Austrian 121,525, Oct 15, 1930 An alloy contg Cu 10-70 and Sb 90-50% is fused with so much Pb and, if desired, unalloyed Sb, that a product contg Pb up to 0.5 and Sb 0.8-1.2% is obtained.

Alloys resistant to acids. SOC. ANON. DES ANCIENS ÉTABLISSEMENTS SKODA A. PLZEN Tr. 37,433 Sept 12, 1929 Addn to 631,578 (C. A. 24, 4254). Alloys of low content of Ti and Si are used.

Alloy for belt conveyors for high-temperature service. WICKWIRE SPENCER STEEL CO Brit. 338,951, Aug 20, 1929 Various structural details are given of belts suitable for use up to temps. of about 870° which may be made of an alloy contg C 0.16-0.24, Cr 17-18, Ni 25-26, Si 2-3 and Fe 55.84-52.76%.

Bearing alloy EMIL ABEL (to Jacob Neurath) U. S. 1,705,633, March 10 In forming a bearing alloy of Pb, Sb, Sn and Ni in which the Ni is present in the α -form, the Ni content is restricted to less than about 1% and Cu is added in an amt. at least 1.5 times that of the Ni content.

Antifriction alloy and its production J. E. ENSLEY and P. STEIN Ifung 101,303, May 14, 1928 In an antifriction alloy contg less than 60% Sn harder Sn-Sb-Cu crystals embedded in the eutectic Cu may be partly replaced by one or more metals of the 3rd group. Also graphite or similar substance can be added in fine dispersion. To the melted Cu or its alloys with the metals of the 3rd group are added first Sn, then Sb and Pb, the liquid mixt. is cooled with stirring, graphite is added; and finally the alloy is solidified.

Aluminum alloys. H. J. GARRARD Brit. 338,923, May 30, 1929 In the production of an Al alloy contg Zn and Mg, with or without Bi, a preliminary alloy of part of the Al with the Zn and Mg is added to the remainder of the Al, and the Bi if used, is subsequently added. Various details of proportions and procedure are given.

Aluminum alloys. O. KAMPS Brit. 339,409, Nov. 25, 1929 Alloys are formed of Al together with Ni 0.3-1.5, Th 0.05-0.9, Mn 0.4-0.8 and up to 6.8% of a Cu-Ce alloy contg 90% Cu. The alloy is quenched from 480-535° to room temp. in a BaCl₂ bath and reheated to 50-300°.

Protecting aluminum alloys. ALADAR PACZ Fr. 695,155, Aug 14, 1929 Al alloys are protected by colored layers obtained by treating the alloys in alk. solns. of chromates or dichromates contg not more than 0.25% by wt. of dichromates calcd. as K₂Cr₂O₇ and at least 1.5% of alk. substance calcd. as Na₂CO₃.

Alloys of iron and aluminum. BRITISH & DOMINIONS FERROALLOY, LTD., and J. W. HAMFYLDF Brit. 338,688, Oct 21, 1929 In making or remelting alloys of Fe and Al, fluorspar or SrF₂ is added. Various details and modifications of procedure are described.

Iron alloys. FERRANTI, LTD. Ger. 515,700, July 11, 1925 An elastic alloy suitable for wire for low frequency windings of telephones consists of Fe 60.5, Ni 25.99, Cr 11.60, W 1.6 and C 0.36%.

Iron alloys hardened by nitridation. H. A. DE FRIES (to Nitralloy, Ltd.) Brit. 338,655, Oct 9, 1928 Alloys suitable for hardening by nitridation (suitably as described in Brit. 174,580, C. A. 16, 1738) contain Fe together with C 0.05-2, Al 2.1-6% and a total of 0.2-8% of one or more of the elements Si, Mn, Ni, Cr, Cu, Mo, W, V, Ti or Zr (the C content being about one third the Al content and the Si content not exceeding the Al content).

Acid-resistant iron alloy. V. F. DONAT and K. V. VLADIMIROV. Russ. appl.

31,146, Aug. 9, 1928. An acid resistant alloy is prepd. from a 70–80% ferrosilicon and a 60–70% ferrochromium and iron scrap proportions to yield about 18% Si and 2–3% Cr in the alloy. Cast iron melted in another furnace is then added to the above alloy.

Iron and its alloys for chemical apparatus. I. G. FARBERND. A-G. Fr. 37,362. Aug. 23, 1929. See Brit. 334,926 (C. A. 25, 1213).

Copper alloys. G. H. WHITEMAN and IMPERIAL CHEMICAL INDUSTRIES, LTD. Brit. 338,676, Oct. 14, 1929. Alloys suitable for tubes contain Cu together with Ni 30 and 4% or less (suitably 2.5%) of Cr. Some mention is also made of the use of larger proportions of Cr.

Copper-aluminum alloys. ALADAR PACZ. Ger. 515,805, Dec. 4, 1927. See Brit. 318,802 (C. A. 24, 2421).

Lead alloys for bearings, etc. METALLGES. A-G. Ger. 518,395, May 19, 1927. Li 0.03–0.05, Na 0.58–0.66, Ca 0.65–0.73, and Al less than 0.2%, the remainder being Pb. Acetylene. I. G. FARBERND. A-G. Ger. 515,916, Aug. 12, 1927. C_2H_2 for autogenous welding is stored and shipped in pressed solid form.

Autogenous welding. I. F. KUBARZHEVSKI. Russ. appl. 49,743, June 21, 1929. An explosive gas mixt. is admitted into the burner simultaneously with acetylene.

Welding composition. JOHN A. HEALY. U. S. 1,796,329, March 17. A compn. suitable for use in welding cutting tips to shanks comprises fused borax 30, ferro-Mn 60, ferro-Si 10 parts, together with high speed steel filings.

Solder. WILHELM REUSS. Fr. 37,159, June 28, 1929. Addn. to 645,549 (C. A. 23, 1866). The solder described in the prior case is used for soldering steel, Fe, Cu, etc.

Treating flux-cored solder. PERRY C. RIPLEY (to Kester Solder Co.) U. S. 1,796,163, March 10. To prevent leakage of flux in storage, solder having a liquid flux core is subjected to a temp. higher than that to which the solder will normally be exposed prior to actual use, to effect an expansion of the liquid flux, and the ends of a strand of the solder are sealed while the liquid flux is in the expanded condition. App. is described.

Soldering flux. DANA W. BOWERS (49% to Charles T. Asbury). U. S. 1,797,124, March 17. Beeswax 6–10, oleic acid 25–40 and Ag oxide 1–3 parts are used with a quantity of finely granulated solder about equal to the other ingredients.

Electrical resistance heating device suitable for use in soldering irons, pipes, etc. EDWIN L. WIEGAND. U. S. 1,795,628, March 10. Structural features.

10—ORGANIC CHEMISTRY

CHAS. A. ROULLER AND CLARENCE J. WEST

Recent advances in science: organic chemistry. J. N. E. DAY. *Science Progress* 25, 389–92 (1931).—A review is given of recent work in org. chem. J. S. H.

Oxidation of methane. I. The homogeneous oxidation of methane under ordinary pressure. KIYOSHI YOSHIKAWA. *Bull. Inst. Phys. Chem. Research* (Tokyo) 10, No. 3, 251–5 (1931) (Abstracts 30–1 (in English), published with *Sci. Papers Inst. Phys. Chem. Research* (Tokyo) 15, Nos. 289–91).— CH_4 is oxidized in the presence of air at 600° to CO , CO_2 , CH_2O and HCO_2H . At 700° H_2 is formed. HNO_3 promotes the start of the reaction to 500°, yielding more CH_2O . $PbEt_4$ has the opposite effect. The reaction is a homogeneous gas reaction of the chain type. ALFRED BURGER.

The reaction of methane and lower homologs with carbon monoxide and dioxide at various catalysts. FRANZ FISCHER AND HERBERT BAHR. *Ges. Abhandl. Kenntnis Kohle* 8, 279–81. *Chem. Zentr.* 1930, I, 31.—A repetition of Berthelot's expt., in which C_2H_6 was formed on conducting 1 vol. CO and 2 vols. CH_4 through a glowing glass tube filled with pumice, yielded the same phenomenon. It was impossible to identify the hydrocarbon on account of the small quantity obtained. In another expt. F. and B. conducted the same mixt. of CH_4 and CO over various catalysts at different temps. CO was decomposed with formation of CO_2 , and CH_4 dissociated partly into C and H, but a reaction between CH_4 and CO was scarcely detectable. When the waste gas from a catalyst for the synthesis of mineral oil was conducted over metallic sponge Fe at 900°, the greatest part of the CO_2 was reduced to CO ; gasol, heavy hydrocarbons and a fraction of H disappeared, and a large quantity of water was formed. G. SCHWOCH.

The decomposition of methane at various temperatures at various catalysts. FRANZ FISCHER AND HERBERT BAHR. *Ges. Abhandl. Kenntnis Kohle* 8, 27–94; *Chem. Zentr.* 1930, I, 31.—When a gas mixt. rich in CH_4 (natural gas from Altengamme with 92.5% CH_4) was conducted through a heated quartz tube or through a heated porcelain

b 60.0–0.4°, d_4^{20} 0.9139, d_4^{25} 0.8961, 81.6%; III, b 53.6–4.0°, d_4^{20} 0.9110, d_4^{25} 0.8915, 88.5%; IV, b 57.8–8.2°, d_4^{20} 0.9121, d_4^{25} 0.8910, 85.2%. V, b 01.2–1.0°, d_4^{20} 0.9081, d_4^{25} 0.8912, 79.8%.

A. I. HENNE

Aliphatic diazo compounds. II. ALEXANDER SCHÖNBURG AND TH. STOLFF (IN PART WITH J. TSCHOLAKOFF AND E. WILLNER) *Ber* 63B, 3102–16 (1930); cf. C. A. 25, 931. CH_3N_2 itself reacts with arylsulfur chlorides like diaryldiazomethanes to give hemimercaptol halides $\text{CH}_3\text{N}_2 + \text{RSX} = \text{CH}_3(\text{SR})\text{X}$ (I) or $[\text{CH}_3(\text{SR})]\text{X} + \text{N}_2$ (II). The aromatic hemimercaptol halides form mercaptols with mercaptans and give free radicals with Hg and other heavy metals. In C.II, they react instantaneously with HgOAg to give $\text{R}_2\text{C}(\text{SR}')\text{OAg}$, the rapidity of the reaction makes it probable that the structure II must be considered along with I. With H_2O they give the corresponding ketone and mercaptol, probably through the hemimercaptol, $\text{R}_2\text{C}(\text{SR}')\text{OH}$. With NiH_2 in C.II, they do not give the expected $\text{R}_2\text{C}(\text{SR}')\text{NiH}_2$ (which is probably an intermediate product) but the mercaptol and ketone imide. It had been found that diaryl thioketones react with 1 mol of certain inorg. metal compds. and it had been concluded that the products are mol compds. (e.g., $\text{R}_2\text{C}=\text{S} \cdot \text{HgCl}_2$). This investigation has been concluded and the results are reported here because these products are now thought to be complexes and very closely related to the aromatic hemimercaptol halides. A systematic study of the action of aromatic thioketones on HgCl_2 , HgBr_2 , AgNO_3 and AgClO_4 has shown that reaction takes place in every case, although the products cannot always be obtained in cryst. form, and sometimes, although cryst., they decompose so rapidly that they cannot be isolated in analytically pure form. The following complex compds., however, were obtained in cryst., analytically pure form: $[\text{R}_2\text{C}(\text{SHgCl})]\text{Cl}$, $\text{R} = o\text{-MeOC}_6\text{H}_4$, red, $p\text{-MeOC}_6\text{H}_4$, orange, $[\text{R}_2\text{C}(\text{SHgBr})]\text{Br}$, $\text{R} = p\text{-MeOC}_6\text{H}_4$, orange, $p\text{-tOC}_6\text{H}_4$, orange, 3,4-MeO($1'\text{O}$) C_6H_3 , orange, $[\text{O}(\text{C}_6\text{H}_5)_2\text{C}(\text{SAg})]\text{ClO}_4$, blood red, $[\text{O}(\text{C}_6\text{H}_5)_2\text{C}(\text{SHgCl})]\text{Cl}$, orange. The color of these compds. is in every case lighter than that of the thioketones from which they are derived, probably because of the disappearance of the C=S group. Some of the compds. are quantitatively converted into the thioketones by covering them with H_2O and shaking out with aq. KI, this is explained by assuming in the Et_2O phase the equil. thioketone + $\text{HgCl}_2 \rightleftharpoons$ complex compd., which is disturbed by the aq. KI. A continuation of the study of the action of diaryldiazomethanes on aromatic disulfides has shown that the formation of mercaptols in this way is also a general reaction, although it is of significance as a method of prepn. only for asym. mercaptols which, for some reason or other, cannot be prepd. by other methods, and in connection with the study of the reaction mechanism. *Diazothioxanthene*, $\text{S}(\text{C}_6\text{H}_5)_2\text{CN}_2$ (I), was obtained from thioxanthone, $\text{S}(\text{C}_6\text{H}_5)_2\text{CS}$ (II), by conversion with N_2H_4 into the hydrazone and oxidation of the latter with HgO or treatment with $\text{S}(\text{C}_6\text{H}_5)_2\text{CCl}_2$. Attempts to prep. diazoxanthene (III) in the same way were unsuccessful, the $\text{O}(\text{C}_6\text{H}_5)_2\text{C}=\text{NNH}_2$ was unchanged by yellow HgO in the cold, and on heating there was obtained the ketazine, $[\text{O}(\text{C}_6\text{H}_5)_2\text{C}=\text{N}-\text{N}-\text{C}_6\text{H}_5]$ (IV), the decompn. product of the intermediate III. IV is also obtained from the hydrazone and $\text{O}(\text{C}_6\text{H}_5)_2\text{CCl}_2$. Chromone hydrazone could likewise not be converted into the diazo compd. (*o*-Nitrophenylmercaptol)chloromethane, from $\text{O}_2\text{NC}_6\text{H}_4\text{SCl}$ and CH_3N_2 in Et_2O , yellow, m. around 95°, shows green halochromism with red fluorescence in concd. H_2SO_4 , the soln. soon becoming brown. *Diphenyl(o-nitrophenylmercaptol)bromomethane*, from $\text{O}_2\text{NC}_6\text{H}_4\text{SBr}$ and PhCN_2 , red-brown, m. about 108°, fumes in the air (IIb) and gradually becomes smeary. *Diphenyl(p-chloro-o-nitrophenylmercaptol)chloromethane*, yellow, m. 121°, evolves HCl on long standing in moist air. *Diphenyl(phenylmercaptol)methyl benzoate*, from $\text{Ph}(\text{PhS})\text{CCl}$ and AgO_2CPh in Et_2O , m. 148–9°, shows red-yellow halochromism in concd. H_2SO_4 . *Diphenyl(o-nitrophenylmercaptol)methyl benzoate* (4.5 g. from 4 g. of the chloride), faintly yellow, m. 132°, shows red orange halochromism in concd. H_2SO_4 . *Diphenylene(phenylmercaptol)methyl benzoate* (2 g. from 4 g. $(\text{C}_6\text{H}_5)_2\text{C}(\text{SPh})\text{Cl}$), faintly yellow, m. around 132°, shows yellow-brown halochromism with concd. H_2SO_4 . *Diphenyl(phenylmercaptol)-(o-nitrophenylmercaptol)methane* (0.1 g. from 5 g. $\text{O}_2\text{NC}_6\text{H}_4\text{SSiPh}_3$ and PhCN_2 refluxed in C_6H_6), yellow, m. 134°. *Di- β -naphthylidiphenylmercaptolmethane*, from $(\text{C}_{10}\text{H}_7)_2\text{CN}_2$ and Ph_2S , m. 160°, shows deep red halochromism and splits off PhSH in concd. H_2SO_4 . The mercaptol could not be isolated from the products of the reaction of (*p*- MeC_6H_4) $_2\text{CN}_2$ with Ph_2S and of Ph_2N_2 with (*p*- MeC_6H_4) $_2\text{S}$ but that it was formed was indicated by the fact that after removal of the ketazine the residual products turned blue at about 210°, and in concd. H_2SO_4 showed halochromism and split off PhSH and $\text{MeC}_6\text{H}_4\text{SH}$, resp. *Xanthone hydrazone*, from xanthone and N_2H_4 , H_2O in abs. alc., light yellow, m. 128–30°, sol. in concd. H_2SO_4 with light yellow color. *Ketazine* (IV), red-yellow,

m 283°, sol in concd H_2SO_4 with orange yellow color. *Thioxanthone hydrazone*, yellowish m 115°, sol in concd H_2SO_4 with orange-red color. I, green, m. 105°, sol in ether, petroleum ether and C_6H_6 with violet color and green fluorescence, in H_2SO_4 with orange red color. *Thioxanthone ketazine*, from the hydrazone with $\text{S}(\text{C}_6\text{H}_5)_2\text{CCl}_2$ in C_6H_6 , yellow red m 281°, sol in H_2SO_4 with bright red color. *Chromone hydrazone*, m 96° was obtained in 4.5 g yield from 5 g chromone with N_2H_4 , H_2O in alc. and in 80% yield from 4 thiochromone. C. A. R.

n-Butyl ester of phosphoric acid. A. ARBUZOV AND I. ARBUZOVA. *J. Russ. Phys. Chem. Soc.* 62, 1533-6(1930).—By acting with PCl_5 on BuONa was obtained tri Bu phosphate $\text{P}(\text{OBu})_3$ (I) identical with the product prepd by Milobedzki and Sachnowski (*C. A.* 13, 2865) by the interaction of BuOH , PCl_5 and $\text{C}_6\text{H}_5\text{N}$. The action of BuI on I gives the isomeric di Bu butanephosphonate, $\text{BuI'O}(\text{OBu})_2$ (II), which when saponid with HCl produces butanephosphonic acid, (III). *Exptl part*—1st was prepd BuONa by heating on the water bath 950 g of freshly distd, *n*- BuOH , b 117° with 70 g of metallic Na then, while cooling, were added Et_2O and PCl_5 , the Et_2O was expelled, and the reaction product on distn gave at 0 mm and 180° (temp of glycerol bath) 174 g (68.6%) of crude I. Close fractionation being impossible, the fraction b 116-30° (97 g) was added to 4.5 g of Na wire suspended in dry Et_2O , then heated 1 hr on the water bath, poured off from unchanged Na, and redistd, producing 27.6% of I b 119.5-20°, d_4^{20} 0.9709, d_4^{25} 0.9201. In the distg flask there remained $\text{P}(\text{OBu})_3\text{ONa}$. When 1.23 g of Br is added to 2 g of I and then distd there is obtained 1.83 g (91.5%) of $\text{C}_6\text{H}_5\text{Br}$, b 98°. The isomerization of I to II is effected on heating 8 hrs at 150-160° in a sealed tube a mixt of 20 g of I and 12 g of freshly distd BuI (b 120-31°), the reaction product on redistg gives 11.50° g of II, b 150-51°, d_4^{20} 0.9674, d_4^{25} 0.9320, colorless and odorless liquid. II is saponid to III on heating 5 hrs at 170° in a sealed tube a mixt of 6.3 g of II and 2 vols of HCl (1:1), the upper layer of BuCl , in theoretical yield, was sepd, and the aq soln was repeatedly evapd with H_2O on a water bath, the heavy mass was treated with H_2O , the sepd crystals filtered off, boiled in H_2O with animal charcoal, filtered and evapd to dryness; yield of crude III 3.55 g, m 80-89°, recrystd from ligroin and dried over H_2SO_4 it m 101-3°, the acid is dibasic. CHAS BLANC

2-Methyl-2-buten-4-ol. F. G. FISCHER, L. EITEL AND K. LÖWENBERG. *Ber.* 64B, 30-4(1931), cl *C. A.* 24, 595.—In connection with the building up of terpene-like substances with open and closed chains it was of interest to prep and characterize 2 methyl 2 buten-4-ol (I). In a note on the rearrangement of tertiary acetylenecarbinols with HCO_2H , Rupe and Kamblis (*C. A.* 20, 3443) state that dimethylethynylcarbinol gives in poor yield "isopropylidenacetaldehyde" (I), the m p of whose semicarbazone is given as 244° in Engl. pat. 267,954. In agreement with them, the authors find that this method is ill adapted to the prepn of I; after short heating of the carbinol with HCO_2H there is obtained a small amount of a C=O compd whose semicarbazone, however, m 221-2° and *p*-nitrophenylhydrazone m 161-2°. After other methods of prepn I from $\text{Me}_3\text{CCHCH}_2\text{CHO}$ had failed, the following method was devised: $\text{iso-AmOH} \rightarrow \text{Me}_3\text{CCHCH}_2\text{CHO} \rightarrow \text{Me}_3\text{CCHCH}_2\text{BrCH}(\text{OEt})_2$ (II) $\rightarrow \text{Me}_3\text{CCHCH}_2\text{CH}(\text{OEt})_2 \rightarrow \text{Me}_3\text{CCHCH}_2\text{CHO}$ (I). The use of fermentation amyl alc as the starting material has the disadvantage that the $\text{Me}_3\text{CCHCH}_2\text{CHO}$ is accompanied by the isomeric EtCHMeCHO , which, however, is removed in the last step of the above reaction it is probably present as $\text{EtC}(\text{CH}_3)_2\text{CHO}$ in the first runnings of the fractionated crude I. The $\text{C}_6\text{H}_5\text{CHO}$, b 90-2°, $[\alpha]_D^{25}$ 5.5° (and therefore contg about 75% $\text{Me}_3\text{CCHCH}_2\text{CHO}$) in CHCl_3 at -25° slowly treated with Br in the light of an Osram lamp and then with abs EtOH gave 75% of the II, b 88-9°, d_4^{20} 1.1772, n_D^{20} 1.4489, which, heated 4 hrs in N with fused KOH yielded 65-70% of the methylbutenal di Et acetal, bn 163-5°, d_4^{20} 0.8555, n_D^{20} 1.4201, decolorizes only about 0.1 of the calcd amt. of Br in CHCl_3 , a slight excess quickly giving a dark green color changing to blood red on addn of NaOH ; Br water is also decolorized but without the play of colors; anhyd $(\text{CO}_2\text{H})_2$ gives a grass-green color and changes the acetal into a viscous liquid converted by beating into a brittle amber like resin insol in the ordinary org solvents but swelling markedly under C_6H_6 . Hydrolysis of 53 g of the acetal with cold satd aq tartaric acid gives 5 g b 26-8°, b 77-80°, of an unsatd. aldehyde, probably $\text{EtC}(\text{CH}_3)_2\text{CHO}$ (semicarbazone, m 183-4°, *p*-nitrophenylhydrazone, golden yellow, m. 141-2°), and 18 g of I, b 62-3°, bn 132-3°, d_4^{20} 0.8722, n_D^{20} 1.4520, is easily auto-oxidizable but stable for a long time in sealed vessels; semicarbazone, m. 221-2°, *p*-nitrophenylhydrazone, violet red, m 161-2°. Hydrogenation of I with Pd-CaCO_3

gives $\text{Me}_2\text{CHCH}_2\text{CHO}$ (semicarbazone, m. 126-7°; *p*-nitrophenylhydrazone, m. 116°)
C. A. R.

Absorption and reactivity of the ketone group. MME. RAMART-LUCAS AND MME. BRUZAU. *Compt. rend.* 192, 427-9(1931)—Variations in the chem. activity of the hexaalkyl-substituted ketones can be explained by the rate of absorption of ultra violet light. This is due to a mutual influence between the alkyl radicals present and the ketone group.
JULIUS WHITE

Catalytic hydrogenation of azines. V. Hydrogenation of ketazines. Determination of affinity capacities of methyl and ethyl groups. K. A. TAIPALE AND P. V. USA. *Chem. J. Russ. Phys.-Chem. Soc.* 62, 1241-58(1930), cf. *C. A.* 20, 3282—The speed of the reaction of catalytic hydrogenation of aldazines makes possible an estimation of the relative consumption of affinity by various radicals at the C atom in the azo methylene grouping of aldazines (*C. A.* 19, 3478). If the affinity capacity of the Me group is 100, the relative values of the affinity capacity of hydrocarbon radicals and the H atom for this C atom can be deduced, as follows: H 12, Me 100, Et 135, Pr 200, iso-Pr 555, iso-Bu 700, Ph 600. Other investigators, detg. the affinity capacity of radicals in relation to C, found that the Me group absorbs a greater amt. of affinity than the Et group. Therefore it is desirable to compare the affinity capacities of Me and Et groups in another example than that previously obtained by hydrogenation of the aldazines of AcH and PrH. The following ketazines are considered: bis(dimethyl)azumethylene, $\text{Me}_2\text{C}=\text{NN}=\text{CMe}_2$ (I), bis(methylethyl)azumethylene, (II), bis(diethyl)azumethylene, (III). From the speed of hydrogenation of I and II, and of II and III, we are able to study the different influences of Me and Et groups on the speed of addn. of H to the azumethylene group $>\text{C}=\text{NN}=\text{C}<$, and from that of I and III the difference in the influence of 2 Me and 2 Et groups, the resulting relations of speeds of hydrogenation enabling us to make an estimate of the relative values of affinity capacities of these 2 radicals. The expts. show that the speeds of hydrogenation of II and I in the presence of Pt black are to each other as 1 to 1.3, those of III and II as 1 to 1.7, and those of III and I as 1 to 2.6, i. e., an azumethylene group joined with Et groups combines with H less readily than the one joined with Me groups. Considering that the speed of hydrogenation of a double bond depends on the magnitude of the residual affinity of the atoms constituting this bond, and that the magnitude of the residual affinity in turn depends on the degree of consumption of the affinity by the radicals combined with these atoms, the conclusion is made that the Et group possesses a higher affinity capacity than the Me group. Having previously assumed the affinity capacity of Me as 100, that of Et then is 140, which corresponds with the value of 135 derived from the speeds of hydrogenation of the aldazines (cf. Kindler, *C. A.* 21, 1453). The catalytic hydrogenation of ketazines led also to investigation of the properties of the fatty hydrazo compds. formed in the reaction. The properties of hydrazo compds. depend on the size and structure of the radicals attached to the hydrazo group, which influence its reactivity, and are noticeable in their basicity, the reactivity of the H of the imino group, and, probably, in the stability of the hydrazo grouping against decomposition into amines by further addn. of H. T. (*C. A.* 18, 3019) has shown that $(\text{NHCHMe}_2)_2$, obtained by hydrogenation of I, possesses a considerably lower basicity than $(\text{NHMe})_2$, $(\text{NHEt})_2$, $(\text{NHPr})_2$ and $(\text{NHCH}_2\text{CHMe}_2)_2$. The $(\text{NHCHMe}_2)_2$ forms a di-HCl salt, which, however, easily loses 1 HCl, moreover, only the H of 1 imino group is capable of reacting, giving mononitro-, mono-Bz derivs. and the corresponding derivs. of $(\text{NH}_2)_2\text{CO}$ and $(\text{NH}_2)_2\text{CS}$, while the other above named hydrazo compds. give disubstituted derivs. Hydrazo- α -methylpropane (IV) and hydrazo- α -ethylpropane (V) obtained, resp. from II and III, behave also like $(\text{NHCHMe}_2)_2$. Their di-HCl salts are even less stable: the stability decreasing with the increasing wt. of the radicals. The more affinity is consumed by the radicals from the hydrazo group, the less of the residual affinity remains with the latter, the lower it seems, is its basicity, and also the lower is the stability possessed by the HCl salts. The hydrazines also form only mono-substitution products with BzCl and PhNCS. IV and V and their homologs are very readily oxidized in the air, forming azo compds., which are isomerized with acids and alkalis to hydrazones, and the latter are decompd. with acids into ketones and primary hydrazines. The hydrogenation of ketazines in Ac_2O proceeds differently than in other solvents. After addn. of 1 mol. of H as in the case of AcOH, the addn. of the next mol. of H is very slow. This is explained from a study of the hydrogenation of II, by assuming that the azo compd. formed in the first phase of the reaction is isomerized to the hydrazone, which is readily acetylated, and the Ac deriv. of the hydrazone under the influence of the Ac group is more difficultly hydrogenated than the free hydrazone or the azo compd. (Skita and Rolles, *C. A.* 14, 3667). *Exptl. part*—The ketazines were prepd. by the known

methods (K and C. C. A. 15, 2879) I b_m 131-3°, b_{11} 38°, d_4^{20} 0.8389, n_D^{20} 1.45318, calcd MR_D 36.14, found MR_D 36.16 II b_m 171-2°, b_{11} 72°, d_4^{20} 0.8101, n_D^{20} 1.4511, calcd MR_D 45.35, found MR_D 45.27 III b_m 196-7°, b_{11} 110.5-11°, b_{11} 82°, d_4^{20} 0.8119, n_D^{20} 1.4574, calcd MR_D 54.58, found MR_D 54.41 For the method of hydrogenation see C. A. 20, 3282. Results of hydrogenation I adds 11.3 times more readily than II, II 1.5 times more readily than III. The products of hydrogenation of the ketazines were isolated from alc. either by drying the filtrate with K_2CO_3 and then redistg. in II, or by satg. the filtrate with HCl , whereby 60-70% of the mono- HCl salt of the hydrazine is pptd. The mother liquor evapd. *in vacuo* deposits the di- HCl salt contaminated with amino- HCl . The hydrogenation in $AcOH$ produces a 100% yield with no formation of amines (C. A. 17, 3015) IV b_m 168.5-9°, b_{11} 100-1°, d_4^{20} 0.8214, d_4^{20} (0.8070) n_D^{20} 1.4283, calcd MR_D 46.23, found MR_D 46.07. The di- HCl salt, loses 1 HCl in the air or H_2O giving the mono- HCl salt, m 147°. Disubutylsemicarbazide, $NH_2CON(C_4H_9)NH(C_4H_9)$, obtained by the action of 1.5 mols of coned. KCN on 1 mol of the HCl salt, m 49-9°. Disobutyl-4-phenylthiosemicarbazide, $PhNHCHN(C_4H_9)NHCH_2$, prepd. by mixing 1 mol of the base with 1 (or 2) mols of $PhNCS$, evapd. and filtering, m 78-85°. Azo- α -methylpropane (VI) was prepd. by adding the calcd. amt of H_2O to 7.5 g. of IV in PtO_2 , while cooling; after 24 hrs. standing the mass was heated on the water bath, filtered, dried with BaO , the Et_2O expelled and redistd., yielding 5.2 g. (70%), b_{11} 140-1°. It is also prepd. by adding 3.5 g. of $NaNO_2$ in H_2O to 6 g. of the HCl salt of IV in H_2O , then a few drops of $AcOH$ is added, allowed to stand 24 hrs., neutralized with Na_2CO_3 , extd. with Et_2O , dried with K_2CO_3 , b_1 31°, d_4^{20} 0.7904, d_4^{20} 0.7749, d_4^{20} 0.7757, n_D^{20} 1.4127, MR_D (calcd.) 45.81, MR_D (found) 45.72. The isomerization of VI to the hydrazone was carried out by refluxing 3 hrs. with $NaOH$, fractionating, taking up the part b 155-64° in Et_2O and satg. with HCl , giving about 50% of isobutylhydrazine 2 HCl , which heated *in vacuo* at 50° goes over to the mono- HCl salt. Dibenzoylisobutylhydrazine, $C_6H_5CH_2NHBz$, prepd. by refluxing several hrs. on the water bath a mixt. of 3 g. of VI in C_6H_5 , 4 g. of $BzCl$ and some calcined Na_2CO_3 , m 169-70°. Isobutylsemicarbazide, $NH_2CO N(C_4H_9)NH_2$, was obtained by heating in a water bath a mixt. of the hydrazine- HCl with KCN in soln., extg. with C_6H_5 , m 97.5-8°. V, obtained by hydrogenation of III, b_m 193-3.5°, d_4^{20} 0.8210, n_D^{20} 1.4358, MR_D (calcd.) 55.51, MR_D (found) 55.16. The di- HCl salt crystallizes with 1 H_2O , the mono- HCl salt is prepd. from the di- HCl salt on warming. Benzoyldiamylhydrazine, $AmNBzNHAm$, was obtained by heating the mono- HCl salt of the base with $BzCl$ and dry Na_2CO_3 in C_6H_5 , b_1 184-5°. Di-amyldsemicarbazide, $NH_2CONAmNHAm$, prepd. by the action of KCN on the HCl salt of the base, m 57-57.5°. Azo- α -ethylpropane, obtained by oxidation of the hydrazo compd. with H_2O in Et_2O , b_m 182°, b_{11} 70°. Amylhydrazine 2 HCl , $Et_3CH-NHNH_2$, 2 HCl , was obtained by the action of HCl on the azo compd. or hydrazone in Et_2O , the mono- HCl salt is formed on heating the di- HCl salt at 60-70°. Dibenzoyl-amyldhydrazine, Et_3CHNH_2Bz , prepd. by heating the azo compd. in Et_2O with an excess of $BzCl$, m 199.5-200°. Amyldsemicarbazide, $NH_2CONAmNH_2$, prepd. from a coned. soln. of HCl -salt of the primary hydrazine on heating with an excess of KCN , m 107-8°.

CHAS. BLANC

A few derivatives of propylene glycol. A DEWAELE. *Bull. soc. chim. Belg.* 39, 395-401 (1930).—The following new compds. have been prepd. $CH_3CH(OH)CH_2OH$ (I) obtained by methylation of $CH_3CH(OH)CH_2OH$ (I) with Me_2SO_4 , b 103-4°, d_4^{20} 1.009, n_D^{20} 1.41372. $CH_3CH(OEt)CH_2OH$, prepd. by ethylation of I with Et_2SO_4 , b 117°, d_4^{20} 0.9965, n_D^{20} 1.41902. The dehydration of I by H_2SO_4 gives $(CH_3CH=CHMe)_2O$ b_m 187-8°, d_4^{20} 1.103, n_D^{20} 1.45046, which reacts with $KOAc$ and $HOAc$ to give $(CHMeCH_2OAc)_2O$, b_m 248°, d_4^{20} 1.050, n_D^{20} 1.42654. $EtOCH_2CH(OH)Me$, obtained from alc. and propylene oxide with 2 drops of H_2SO_4 , b 136°, d_4^{20} 0.9028, n_D^{20} 1.4. $EtOCH_2CHMeOAc$, b 158-60°, d_4^{20} 0.9461, n_D^{20} 1.40968. $CH_3(OH)CH(OEt)Me$, obtained from I through the acetin, b_m 140-1°, d_4^{20} 0.9044, n_D^{20} 1.40966, is obtained from $MeOH$ and propylene oxide, with H_2SO_4 . $MeOCH_2CH(OAc)Me$, b_m 147°, d_4^{20} 0.9709, n_D^{20} 1.40449. $Me_2CHOCH_2CH(OH)Me$, b_m 142-3°, d_4^{20} 0.9059. The condensation of $HOCH_2CHClMe$ and propylene oxide with H_2SO_4 gives $MeCHClCH_2OCH_2CH(OH)Me$, b_m 204-5°, d_4^{20} 1.059. $CH_3(OCOPr)CHClMe$, b_m 184°, d_4^{20} 1.031, n_D^{20} 1.42875. $AcOCH_2CH(OAc)Me$

b_{11} 190–1°, d_{20}^{25} 1.059, n_D^{20} 1.4173, treated with HCl gives a mixt. of $\text{ClCH}_2\text{CH}(\text{OH})\text{Me}$ (preponderant) and $\text{ClH}_2(\text{OH})\text{CHClMe}$ while HBr gives a mixt. of the 2 corresponding Br derivs

A. L. HENVE

Ether-like compounds. III. Polyether acids of the type $\text{R.O.}(\text{CH}_2\text{CH}_2\text{O})_n\text{CH}_2\text{CO}_2\text{H}$. M. H. PALOMAA AND T. A. SITTÖNEN. *Ber* 63B, 3117–20 (1930); cf. *C. A.* 13, 2863.—The mono- and di-ether acids of the above type ($n = 1$ and 2) are readily obtained by converting the com "alkylglycols" and "alkylpolyglycols" into the corresponding Na alcoholates, adding $\text{ClCH}_2\text{CO}_2\text{H}$ and liberating the acids from the resulting Na salts. The Na alcoholates are very sol. in ether or the corresponding ether alc and on evapn of the solvent yield crystals which have not yet been investigated. The poly-ether acids are also relatively easily sol in H_2O and their extn. (especially of the MeO derivs) from H_2O with ether is a time-consuming operation. The MeO compds exchange the Me for H with HI (β -Methoxyethoxy)acetic acid (yield, 80%, based on the $\text{ClCH}_2\text{CO}_2\text{H}$), thick oil miscible with water in all proportions, b_{11} 121–2°, d_{20}^{25} 1.1634, n_D^{20} 1.43375, 1.43592, 1.44129, 1.44556 at 20° for α , D, β and γ ; Et ester, from the acid in abs alc with CuSO_4 and a few drops concd H_2SO_4 , miscible with water in all proportions, b_{11} 90°, d_{20}^{25} 1.0369, n_D^{20} 1.41697, 1.41908, 1.42412, 1.42836 (β Ethoxyethoxy)acetic acid, b_{11} 125–6°, d_{20}^{25} 1.1103, n_D^{20} 1.43349, 1.43572, 1.44100, 1.44534 (β -Propoxyethoxy)acetic acid (from ethylenglycol mono-Pr ether, b_{11} 150–0.5°, d_{20}^{25} 0.9106), b_{11} 131°, d_{20}^{25} 1.07415, n_D^{20} 1.43432, 1.43652, 1.44186, 1.44604 (β -Butoxyethoxy)acetic acid (from redistd com "butylglycol," b_{11} 168–9°, b_{11} 50°, d_{20}^{25} 0.9015, n_D^{20} 1.41711, 1.41980, 1.42493, 1.42901), miscible with water, b_{11} 141°, d_{20}^{25} 1.04635, n_D^{20} 1.43605, 1.43823, 1.44360, 1.44797. Repeatedly refractionated $\text{MeOCH}_2\text{CH}_2\text{OCH}_2\text{CH}_2\text{OH}$ (com "methyl polyglycol"), b_{11} 65°, d_{20}^{25} 1.02695, 1.42475, 1.42686, 1.43197, 1.43620, gave β -(β -methoxyethoxy)ethoxy)acetic acid, miscible with water, b_{11} 155–6°, d_{20}^{25} 1.1492, n_D^{20} 1.44351, 1.44575, 1.45112, 1.45553

C. A. R.

Itaconic anhydride and itaconic acid. R. L. SHRINER, S. G. FORD AND L. J. ROLL. *Org. Syntheses* 11, 70–2 (1931).—Heating citric acid until it melts and then rapidly distg gives 37–47% of itaconic anhydride, the free acid is obtained by refluxing with H_2O for 1 hr.

C. J. WEST

Resolution of synthetically prepared isoleucine into its four optically active components, namely *l*- and *d*-isoleucine and *d*- and *l*-alloisoleucine. EMIL ABDERHALDEN AND WALTHER ZEISSET. *Z. physiol. Chem.* 195, 121–31 (1931).—Isoleucine, $\text{EtCHMeCH}(\text{NH}_2)\text{CO}_2\text{H}$, contains 2 asym. carbons, and its synthetic prepn. should therefore consist of 2 pairs of optical isomers. The 2nd pair has hitherto been overlooked. Isoleucine was prepd by the previously described method (*C. A.* 4, 172) in which $\text{EtCHMeCHBrCO}_2\text{H}$ is aminated with NH_4OH . The sepn. of the isomeric pairs was effected by repeated ppto of the aq soln. by EtOH in which *dl*-isoleucine is insol., and recovery of *dl*-alloisoleucine from the mother liquor by distn. of the Et ester and sapon., of the distillate by boiling with H_2O . Sepn. of the optical isomers was effected by converting the *dl*-mixt. into the formyl deriv. and crystg the brucine salt of the latter. The properties of the 4 individual isomers and their HCO , PhSO_2 , PhNCO and $\alpha\text{-C}_6\text{H}_5\text{NCO}$ derivs are described. *l*-Isoleucine m 285–6° (decompn.), $[\alpha]_D^{25}$ in H_2O 10.7°, in 20% HCl 40.8°; *l*-alloisoleucine m 278° (decompn.), $[\alpha]_D^{25}$ in H_2O 14.0°, in 20% HCl 38.1°; *d*-isoleucine, m 283–4° (decompn.), $[\alpha]_D^{25}$ in H_2O –10.7°, in 20% HCl –41.6°; *d*-alloisoleucine, m 274–5° (decompn.), $[\alpha]_D^{25}$ in H_2O –14.2°, in 20% HCl –38.0°; formyl-*l*-isoleucine, m 155°, $[\alpha]_D^{25}$ in EtOH 26.6°; formyl-*l*-alloisoleucine m 126°, $[\alpha]_D^{25}$ in EtOH 24.2°, formyl-*d*-isoleucine, m 156°, $[\alpha]_D^{25}$ in EtOH –26.8°; formyl-*d*-alloisoleucine, m 126°, $[\alpha]_D^{25}$ in EtOH –25.2°; benzenesulfonyl-*l*-isoleucine m 153°, $[\alpha]_D^{25}$ in N NaOH –14.4°, in EtOH 25.3°; benzenesulfonyl-*l*-alloisoleucine, m 147–8°, $[\alpha]_D^{25}$ in EtOH 30.7°; benzenesulfonyl-*d*-isoleucine, m 153–4°, $[\alpha]_D^{25}$ in N NaOH 14.3°, in EtOH –25.5°; benzenesulfonyl-*d*-alloisoleucine, m 147–8°, $[\alpha]_D^{25}$ in EtOH –30.7°; *l*-isoleucine phenyl isocyanate deriv., m 121°, $[\alpha]_D^{25}$ in N NaOH 14.9°, in EtOH 37.5°; *l*-alloisoleucine phenyl isocyanate deriv., m 151°, $[\alpha]_D^{25}$ in NaOH 16.9°, in EtOH 30.8°; *d*-isoleucine phenyl isocyanate deriv., m 119–21°, $[\alpha]_D^{25}$ in NaOH –15.0°, in EtOH –36.3°; *d*-alloisoleucine phenyl isocyanate deriv., m 151°, $[\alpha]_D^{25}$ in NaOH –16.8°, in EtOH –30.6°; *l*-isoleucine α -naphthyl isocyanate deriv., m 178–9°, $[\alpha]_D^{25}$ in EtOH 30.1°; *l*-alloisoleucine α -naphthyl isocyanate deriv., m 165–6°, $[\alpha]_D^{25}$

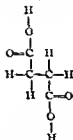
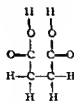
γ -EtOH 25.1°. *d*- α -isoleucine α -naphthyl isocyanate deriv., m. 177-8°, $[\alpha]_D^{25}$ in EtOH +20.5°. *d*- α -isoleucine α -naphthyl isocyanate deriv., m. 168°, $[\alpha]_D^{25}$ in EtOH -25.5°. *trans*-*dl*-isoleucine m. 120-1°, *formyl*-*dl*-alloisoleucine, m. 117-8°; and *dl*-alloisoleucine Et ester b_p 53-8° (7-8° lower than the recorded b_p of *dl*-isoleucine Et ester).

A. W. DOX

Action of thionyl chloride on some anilic acids. W. H. WARREN AND R. A. BAIGGS. *Ber* 64B, 20-30 (1931).—By the action of PCl_5 on $\text{HO}_2\text{CCONHPh}$ Achen obtained ClCOCONHPh which on heating readily lost HCl and CO and yielded about 85% PhNCO . He suggested that oxazul (I) was possibly an intermediate product, but the properties of I are so characteristic that it could not have escaped detection if it had been formed; moreover, it is not affected by PCl_5 . In attempting to prep PhNCO by this method but with SOCl_2 instead of PCl_5 , W. and B. obtained I instead of PhNCO , and they found that a no. of other anilic acids react in the same way with SOCl_2 . The anils were obtained by refluxing the anilic acid with 8 parts SOCl_2 in an all glass app until the evolution of HCl and SO_2 ceased, distg off the excess of SOCl_2 , washing with petroleum ether (dried over Na) or CaH_2 , extg the last traces of anilic acid with boiling H_2O and crystg from an appropriate solvent. I, crystd. from PhNO_2 , does not m. 320°, mol wt in boiling PhNO_2 154.4, unchanged by most dil or concd, hot or cold mineral acids but sol in hot concd H_2SO_4 with decompn into PhNH_2 , CO and CO_2 , decompd by hot NaOH with evolution of a PhNC odor, converted by boiling $\text{Ba}(\text{OH})_2$ into PhNH_2 and $(\text{CO})_2\text{Ba}$, by alc. NH_3 in sealed tubes into $(\text{CONHPh})_2$, and $(\text{CONH}_2)_2$, and by excess of boiling PhNH_2 into $(\text{CONHPh})_2$. *Malonanil*, light yellow, m. 249°, similarly obtained from $\text{HO}_2\text{CCH}_2\text{CONHPh}$, is apparently not affected by concd HNO_3 or HCl , decompd by hot concd H_2SO_4 into PhNH_2 , CO_2 and AcOH , converted by alc. NH_3 in sealed tubes into $\text{CH}_3(\text{CONHPh})_2$ and $\text{CH}_3(\text{CONH}_2)_2$, and by boiling NaOH into PhNH_2 and $\text{CH}_3(\text{CO-Na})_2$. In the same way were prepd succinanil, m. 150°, glutaranil, does not m. 320° (possibly a polymer of Bödtker's compd, m. 144-5°, obtained directly from glutaric acid and PhNH_2 (*Diss. Leipzig*, 1871)), phthalanil, m. 203°, camphanil, m. 116°, and diphenanil, m. 109° (*diphenanilic acid*, from diphenic anhydride and PhNH_2 , m. 170°).

C. A. R.

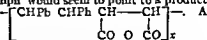
The polymorphism of the saturated diacids of the aliphatic series as a function of temperature. F. DURRÉ LA TOUR. *Compt rend* 191, 1348-50 (1930).—Two forms of succinic acid have been studied under the microscope and by x rays. β -Succinic acid, stable at low temps., is trichinic pinacoidal with a 5.04 Å U., b 7.53 Å U., c 9.75 Å U. At about 137° a transformation to the α -form occurs. The α -acid is monoclinic prismatic with a 5.03 Å U., b 8.75 Å U., c 7.45 Å U. The mol structure appears to be modified thus in the cryst. change

Mol β (less stable)Mol α

H. W. WALKER

Additive heteropolymerization. THEODOR WAGNER-JAUREGG. *Ber* 63B, 3213-24 (1930).—In the investigation of highly polymerized org. natural products Staudinger and his pupils have made use of model compds. prepd by polymerization, or addn to each other, of several identical mols. of low mol wt. It is now shown that 2 unlike units of low mol wt., each contg a C-C union, can likewise be combined into a large mol by polymerizing addn. The name *additive heteropolymerization* is suggested for such a process. The present paper deals with the prepn. of addn heteropolymerizates without the use of catalysts from 2 ethylene components, 1 of which was maleic anhydride (I). I alone shows no tendency to polymerize but has a pronounced additive capacity. When it is boiled with stilbene (II) in xylene, there seps. 45-70% of a white, amorphous, infusible substance (III) insol in most org. solvents, sol colloiddly, with preliminary

swelling, in others (including camphor). Because of its slight soly., its mol. wt. could not be detd., but it is doubtless very high, for the accompanying polymer homologs extd. in very small amount by boiling xylene have in molten camphor an av. particle size of about 4200, corresponding to an av. content of 15 mols. each of I and II. On dry distn. with or without $\text{Ba}(\text{OH})_2$, it decomps. with regeneration of II. From its compn. it consists very nearly of equal nos. of I and II mols., irrespective of the relative proportions of the 2 components in the original reaction mixt. (the ratio II:I in the product obtained from a 2:1 mixt. is 1:1, and in that obtained from a 1:2 mixt. it is 0.95:1). The constancy of the stoichiometric compn. would seem to point to a product built up according to a definite plan, such as—



soln. of the product in $(\text{CHCl}_3)_2$ absorbed Br only after long boiling and in very small amount. Along with the III there is formed a very small amount (about 10% of the wt. of the III) of very slightly more sol. polymerization products, the rest of the original materials remains unchanged. The polymerization does not take place in the lower-boiling C_6H_6 and PhMe , but there also seems to be a sp. solvent factor, for the yield is much smaller in $(\text{C}_6\text{H}_5)_2$, which boils about 6° higher than xylene and even less in the molten mixt. at the temp. of boiling xylene. Conversely, benzalfluorene (IV) and I give in good yield in the fused mixt. at 130° , but not in xylene, a white amorphous heteropolymer very similar to III, insol. in the fused mixt. of the components; no intermediate products of lower mol. wt. were observed; the ratio IV:I is 1:1 and 0.9:1, resp., in the products obtained from 2:1 and 1:2 mixts. of IV and I. The fused mixt. of anisalfluorene and I at about 130° also deposits an amorphous polymer. The heteropolymers of $\text{PhCH}_2\text{CH}_2\text{CH}(\text{Ph})$ (V) and I are likewise sol. in the fused mixt. of the monomers. The ratio V:I is not 1:1 but shifted in favor of I (1:1.5 and 1:1.3–1.25), the V are relatively easily sol. Attempts to prep. heteropolymerizates of bromostilbene, $\text{Ph}_2\text{C}(\text{CPh}_2)_2$, $(\text{PhCH}_2)_2\text{C}(\text{Ph})_2$, and other unsatd. compds. with I were unsuccessful, indicating that the nature of the substituents on the C—C union has an influence on the tendency to heteropolymerization. Steric influences also play a role, for isostilbene with I gave a polymer in only very poor yield. While the hydrocarbons which are well adapted to heteropolymerization show no marked tendency to homopolymerization, the easily polymerized $\text{Ph}_2\text{C}(\text{CH}_3)_2$ forms with 2 mols. I a well crystd. compd. (VI) having the calcd. mol. wt. 376. Succinic anhydride cannot be substituted for I in these reactions but $(\text{PhCH}_2)_2$ gives, although in very poor yield, a 1:1 heteropolymerizate with I. Amylene and styrene heated with I give viscous, gummy masses. Styrene and di-Me maleate gave a white amorphous powder which, from its compn., contained 4–5 styrene mols. to each ester mol. and gave di-Me fumarate on dry distn. Some substances with a C—N union also seem to form heteropolymerizates with I; I boiled with fluorenone imine in C_6H_6 or with benzal- or cinnamalazine, in PhMe or xylene gives yellow to brown, amorphous products; benzalazine, in addn., gives a cryst. addn. product (VII) contg. 2 mols. I, and, with di-Me maleate in xylene, an analogous product (VIII). VI, $\text{C}_{22}\text{H}_{18}\text{O}_4$, m. $279\text{--}81^\circ$ (cor., decompn.), mol. wt. in camphor $373\text{--}8$, sol. in hot 2 N NaOH and boiling 2 N Na_2CO_3 gives a green color with cold alk. KMnO_4 decolorizes KMnO_4 in Me_2CO only very slowly, does not decolorize Br in boiling AcOH. VII, $\text{C}_{24}\text{H}_{18}\text{O}_5\text{N}_2$, m. $243\text{--}4^\circ$ (cor.) VIII, $\text{C}_{24}\text{H}_{18}\text{O}_5\text{N}_2$, m. $232\text{--}4^\circ$ (cor.).

C. A. R.

Glucomannan from "konjak." I. KITSUJI NISHIDA AND HIDEO HATOGI. *J. Agr. Chem. Soc. Japan* 6, 991–5, *Bull. Agr. Chem. Soc. Japan* 6, 78–80 (1930).—The mannan was liquefied by boiling with H_2O or by enzymes and pptd. with alc. The mannan still contains the other impurities such as inulin, pectin, galactan, fructan, xylan, araban, etc. The powd. crude mannan was beaten with H_2O under 1–1.5 atm. in an autoclave. It was pptd. by Fehling soln. and decompd. by alc. HCl. Pure mannan was pptd. by alc. and Et_2O . It is a white powder which has no 1 reaction and does not reduce Fehling soln. Its ash content is 0.10–0.45%. II. *J. Agr. Chem. Soc. Japan* 6, 1065–77; *Bull. Agr. Chem. Soc. Japan* 6, 80–2 (1930).—The mannan was hydrolyzed by mineral acids of various concns. The hydrolysis was complete with 2% HCl after 25 hrs. The ratio of mannose to glucose was 1:1 from the detn. of mannose as the phenylhydrazones and 2:1 by calcn. from the rotation. The difference is due to the incomplete pptn. of mannose phenylhydrazone. When a large amt. of PhNHNH_2 was added, the ratio was 2:1. Glucomannan of "konjak" consists of 2 mols. of mannose and 1 mol. of glucose.

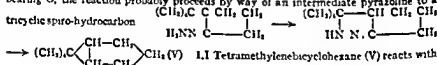
Y. KIHARA

A mannan from the bulb of *Lilium*. I. TEIZO TAKAHASHI. *J. Agr. Chem. Soc. Japan* 6, 791–9 (1930).—The water ext. of the hly bulb powder was filtered and

pptd with alc. The ppt was hydrolyzed by dil HCl. Only mannose and glucose were identified, resp., as phenylhydrazone and osazone in the hydrolyzate. The mannan obtained was named "yurl mannan". Tri Ac deriv. is a white powder, sol. in AcOH, CH_2Cl_2 , acetone, AcOH, and pyridine, m p about 215° , $[\alpha]_D^{25} -17.61^\circ$. II. Ibid. vol. 9.—The mannose and glucose ratio is 2:1. The aldose detn. method of Willstätter and Schindler was extd. for mannose, galactose, glucose, fructose, arabinose and xylose. When 3.5–5 times the 0.1 N I theoretically required was used, mannose was detd. quant. In "yurl mannan" no ketose is present. Y. KIHARA.

The dispersion of the refraction of cyclic hydrocarbons. MARCEL GODCHOT AND (Mlle.) G. CAUQUIL. *Compt. rend.* 191, 1326–8 (1930).—From n_D^{15} and n_D^{25} and d_{15}^{25} the sp. dispersion $\Delta n/d$ has been calcd. for a no. of satd. and unsatd. cyclic hydrocarbons of C_3 , C_4 , C_5 , and C_6 . Δn increases with increasing mol. wt. and the value for the unsatd. compds. is more than 20 units greater than that for the corresponding satd. compd. $\Delta n/d$ has the nearly const. value of 121×10^{-4} for the satd. compds. and 151×10^{-4} for the unsatd. compds. It is independent of the nature of the ring, being the same for methylenecyclopentane, cyclohexane and methylenecyclooctane and likewise the same for cyclooctene and methylenecycloheptene. Aromatic hydrocarbons with $\Delta n/d$ about 250×10^{-4} may be detd. in oils in the presence of cyclic hydrocarbons. II. W. WALKER.

Cyclopentylidenecyclopentane and its relation to catalytic hydrogenation. N. D. ZELINSKI AND N. I. SURIKIN. *J. Russ. Phys. Chem. Soc.* 62, 1313–54 (1930), cf. C. A. 25, 1513.—After the failure to obtain 1,1-dihydroxydicyclopentyl (pinacol of cyclopentanone) for the prepn. of spirocyclodecane (C. A. 24, 76), the attempts to obtain the pinacol from cyclopentylidenecyclopentane (I) by oxidation with BrO_3H , or to convert I into the glycol with KMnO_4 , have also met with negative results. The hydrogenation of I in presence of Pt C produced a satd. hydrocarbon different from the expected dicyclopentyl (II). These 2 negative results have led to an investigation of the starting product, cyclopentylidenecyclopentanone (III). III was prepd. by Wallach's method (Ber. 29, 2755 (1896)), and by dehydrating cyclopentanone (IV) vapors with Al_2O_3 at 200° . Of interest is a large increment in the mol. refraction of III, while I derived from III shows considerable optical decrement. The high coeff. of refraction of III may indicate that a considerable part of III is present in tautomeric (enol) form. The relation of the structures of III and I may be elucidated by decoupling of the hydration of III in the presence of Pt, whereby instead of substitution of H_2 for N at the C bearing O, the reaction probably proceeds by way of an intermediate pyrazoline to a



Br and KMnO_4 as an unsatd. compd., and with BrO_3H not at all. By catalytic reduction must be formed a satd. hydrocarbon $\text{C}_{10}\text{H}_{18}$, spirocyclodecane, $(\text{CH}_2)_5\text{C} \text{---} (\text{CH}_2)_5\text{CH}_2$ (VI), which evidently undergoes isomerization to methylspirocyclononane, $(\text{CH}_2)_5\text{C} \text{---} \text{CH}(\text{Me}) \text{CH}_2 \text{CH}_2 \text{CH}_2 \text{CH}_2 \text{CH}_2$ (VII). The phys. properties of V are almost identical

with those of IV, and greatly different from those of II. *Exptl. part*.—The prepn. of III from IV with such agents as alc. EtONa (Wallach), solid KOH, alc. KOH, or KCN gives up to 40% of higher condensation by products, while passing vapors of IV over Al_2O_3 at 200° results in a 41% yield of III with practically no formation of by-products. A mixt. of IV and concd. alc. KOH is allowed to cool, then heated 1.5 hrs. on the water bath, treated with H_2O , the oily layer extd. with Et_2O , dried with solid KOH, the Et_2O expelled, and distd. in vacuo, $b_{16} 116\text{--}8^\circ$, $n_D^{20} 1.5211$, $d_4^{20} 1.0172$, oxime m $122\text{--}3^\circ$, semicarbazone, m $167\text{--}70^\circ$ (decomps.), hydrazone, m $88\text{--}04^\circ$ (decomps.). III obtained by heating IV 4 hrs. with 10% solid KCN, $n_D^{20} 1.5186$, $d_4^{20} 1.0172$, $b_{16} 119\text{--}21^\circ$, $b 252\text{--}4^\circ$. The fraction $b_{16} 185\text{--}200^\circ$ on recrystn. from EtOH m $83\text{--}4^\circ$ and analyzes for a compd. to which Wallach ascribes the m p $76\text{--}7^\circ$, and the formula $(\text{CH}_2)_5\text{C} \text{---} \text{C} \text{CO} \text{C} \text{---} \text{C}(\text{CH}_3)_2$. V, obtained by distg. the hydrazone of III in



presence of KOH and Pt, treating with dil. AcOH, washing with H_2O to a neutral reaction, drying with K_2CO_3 , and twice distg. over metallic Na, colorless liquid, reacting with KMnO_4 and Br, $b_{16} 189\text{--}90^\circ$, $n_D^{17} 1.4864$, $d_4^{17} 0.9131$. V produces with KMnO_4

no pinacol, and does not react with $\text{Br}_2/\text{H}_2\text{O}$. By passing V 3 times over Pt-C at 190° and subsequently distg over Na is obtained VII, b_m , $184-5^\circ$, n_D^{20} 1.4708, n_D^{25} 1.4701, d_4^{20} 0.8728, which reacts with KMnO_4 very slowly, and on dehydrogenation over Pt-C at 300° shows no change. 1-Cyclopentenylcyclopentane, (VIII), was obtained by dehydrating cyclopentyl 2-cyclopentanol (IX), prepd by reduction of III with Na in aq Et_2O . VIII, b_m , $180-91^\circ$, n_D^{20} 1.4853, d_4^{20} 0.8898, reacts with KMnO_4 and Br, and when reduced over Pt-C at 190° gives II. Cyclopentyl 2-cyclopentanone (X), obtained when 33 g of IX is mixed with 80 cc of 80% AcOH , to which is added 17 g of CrO_3 in 80 cc of H_2O , the whole is heated 30 min on the water bath, dild with H_2O , the oily layer sepd, washed with H_2O , dried with calcined K_2CO_3 , and redistd, b_m , $111-3^\circ$, semicarbazone, m 210° , d_4^{21} 0.9748, n_D^{22} 1.4793. *Prepn of II.* The hydrazone was prepd, by heating on the water bath 23 g of X with 11.5 g of N_2H_4 , H_2O in alc., the crude product was decompd with KOH in presence of Pt powder, dried with CaCl_2 and redistd over Na, b_m , $188-89.5^\circ$, d_4^{21} 0.8652, n_D^{21} 1.4638. CHAS BLANC

A hydrocarbon from pine oil essence and the stereoisomeric 1,2-dimethylcyclopentanes, G CHAVANNE *Bull soc chim Belg* 39, 402-11(1930).—A fractional distn of $\text{C}_{11}\text{H}_{22}$, prepd by dehydration of 1,2-dimethyl-1-cyclopentanol, gives about

10% of $\text{MeC CMe CH}_2\text{ CII}_2\text{ CII}_2$, b $96-7^\circ$, d_4^{20} 0.783, n_D^{20} 1.4321, n_D^{25} 1.4415, n_D^{30}

1.4472, n_{Br} 1.4349, m -124.8° , and 90% of $\text{HC CMe CHMe CH}_2\text{ CII}_2$, b $105.0-5.2^\circ$, d_4^{20} 0.7952, d_4^{25} 0.7998, n_D^{20} 1.4412, n_D^{25} 1.4511, n_D^{30} 1.4571, n_{Br} 1.4442, viscosity η_{10} 0.00537, m -90.4° . These synthetic hydrocarbons are compared with $\text{C}_{11}\text{H}_{22}$ from pine oil, and it is concluded that the natural product is probably the isomer b 105° , with a small amount of impurity which affects its d and n, but not its b p. The hydrogenation of $\text{C}_{11}\text{H}_{22}$ with Pt black gives the 2 stereoisomeric forms of 1,2-dimethylcyclopentane, one of which b $91.8^\circ \pm 0.1^\circ$, b_m 90.95° , d_4^{20} 0.7495, d_4^{25} 0.7541, m -120° , n_D^{20} 1.4093, n_D^{25} 1.41659, n_D^{30} 1.42077, n_{Br} 1.41155, while the other b $99.25^\circ \pm 0.1^\circ$, b_m 93.40° , m -62° , d_4^{20} 0.7718, d_4^{25} 0.7764, n_D^{20} 1.42014, n_D^{25} 1.42748, n_D^{30} 1.43180, n_{Br} 1.42202. The spatial configuration is discussed. A. L. HEVNE

β -Enines and β -dienes V. GRIGNARD AND L. LAPAYRE *Compt rend* 192, 250-3(1931).—The influence of 2 triple bonds, or 1 triple and 1 double bond, in the β position, on the acidity of the central CII_2 group was investigated. For this purpose the reactions of 1,5-diphenyl 1,4-pentadiene $\text{PhC CCH}_2\text{C CPh}$ (I), 1-phenylpent-4-en-1-ine, $\text{PhC CCH}_2\text{C CII}_2$ (II), and dec-1-en-4-ine $\text{Me(CCH}_2)_2\text{C CCH}_2\text{CII}_2$ (III) with Na, NaNH_2 , EtMgBr were studied. By heating I with EtMgBr , 1 mol of $\text{C}_{11}\text{H}_{14}$ is evolved at 35° , 2 mols at 70° . II and III show monoacidity at 70° , III does not evolve 2 complete mol equivs of $\text{C}_{11}\text{H}_{14}$ at 175° . NaNH_2 evolves more than one NH_2 from I and III, in consequence of a partial rearrangement to derive $\text{C}_{11}\text{H}_{14}$ with free active H. The triple bond exerts a more negativating effect on the CII_2 group in these systems than the double bond, the aromatic substituents reinforce this action. I, prepd from CH_2I_2 and Mg phenylacetylide (IV), m $83-90^\circ$, yield 8%. Allyl bromide (V) and IV yield 70% of II, a liquid, b_m , $103-5^\circ$, d_4^{15} 0.940, n_D^{18} 1.558, exaltation of the mol refraction 1.9. Mg phenylacetylide and V yield III, a liquid b_m , $82-5^\circ$, d_4^{11} 0.806, n_D^{15} 1.459, exaltation 0.4. It becomes resinous on standing. ALFRED BURGER

Sulfides and polysulfides of organic bases T. G. LEVI *Gazz chim ital.* 60, 975-87(1930).—A review of the literature shows a great scarcity of expts on the prepn. of polysulfides of org bases. In view of this, several polysulfides were prepd, and their properties studied. Phenylbiguanide- H_2S (I) was prepd by the method of Romani (cf C. A. 18, 1066), and its analysis, not undertaken by R., confirmed its compn. Phenylbiguanide sulfide could not be obtained as indicated by R. (who did not analyze it), i. e., by slow conversion of I in air, or rapid conversion on heating, with loss of H_2S ; and even by the action of concd aq (NH_4) $_2\text{S}$ on phenylbiguanide only I is formed, and on the other hand when aq I is heated gently, H_2S is eliminated, but no neutral sulfide is formed. All attempts at the prepn of diphenylbiguanidine sulfide were without success. Alc. o-tolylbiguanide treated with H_2S , and Li_2O added, ppts in good yield o-tolylbiguanidine hydrosulfide $\text{C}_{11}\text{H}_{13}\text{N}_3$, H_2S , m $123-6^\circ$, stable at room temp when kept dry, decomps slowly in cold water. Benzylamine polysulfide, $(\text{PhCH}_2\text{NH}_2)_n$, H_2S , (II), was prepd in a different way from that described by Bruni and Borgo (cf C. A. 2, 2659). Addn of S to PhCH_2NH_2 (III) gives the orange color characteristic of poly-

sulfides, if at the same time H_2S is passed in, there is an abundant ppt. of II, but it is so unstable that it cannot be washed and dried to const. wt. without loss of H_2S . When dil aq III HCl is poured into excess NH_4 polysulfide (approx compn $(\text{NH}_4)_2\text{S}_x$ prepd from NH_4OH of d 0.92), II is pptd and under these conditions it is more stable, and it can be washed with water without losing H_2S , but when it is dried in a desiccator or in air, it decomps slowly. Since it was impracticable to obtain II dry enough to analyze, it was analyzed in the wet state. This analysis indicated $(\text{PhC}_6\text{H}_4\text{NH})_2\text{H}_2\text{S}_x$. Further attempts to isolate polysulfides of other aliphatic and heterocyclic bases particularly secondary bases, were unsuccessful, though the formation of the polysulfides was proved. When S is added to $\text{C}_6\text{H}_5\text{NH}$ (IV) the liquid turns more intensely orange than with III, and heat is evolved (cf. Bedford, U. S. Pat. 1,710,920 (C. A. 21, 4376)). To isolate the IV polysulfide (V), the reaction product was washed with Et_2O and with CS_2 , but during this it decomps with evolution of H_2S . The reaction mixt was therefore extd. with cold water, but the yellow water ext. let stand rapidly became turbid and S sepd. With acids it forms II persulfide, which proves the formation of V. V is readily decompd by excess PbO , in which case the liquid becomes colorless, and IV probably forms a polythioamine. By filtering off the PbS , spontaneous evapn, soln in EtOH and partial pptn with water of the residue from the evapn, the polythiopiperidine seps, and can be purified by recrystn from EtOH . The reactions are (1) $2\text{IV} + x\text{S} \rightarrow (\text{C}_6\text{H}_5\text{N})_2\text{S}_{x+1} + \text{H}_2\text{S}$, and (2) $2\text{IV} + \text{H}_2\text{S} + x\text{S} \rightarrow (\text{C}_6\text{H}_5\text{N})_2\text{H}_2\text{S}_x$. Data on N polythioamines higher than the dithio derivs. are to be published in a forthcoming paper. Et_3NH and S behave similarly at ordinary temp. With all these bases, attempts to sep. the polysulfides from solns. of salts of the corresponding bases and $(\text{NH}_4)_2\text{S}_x$ were not successful. The method of Romani (loc. cit.) for the prepn. of phenylbiguanide polysulfide, $(\text{C}_6\text{H}_5\text{N})_2\text{H}_2\text{S}_x$, viz. by boiling aq I with S, was not considered satisfactory, because, as was verified by expt., there is too great a loss of H_2S and sepn. of phenylbiguanide (VI). Accordingly concd alc. VI contg S was treated with H_2S , under which conditions a mixt. of I and a polysulfide (VII) pptd., and the latter alone could not be obtained because I is insol and seps immediately, and then is transformed slowly and incompletely into VII. In contrast to this, $(\text{PhIN})_2\text{C NH}$ does not form directly a sulfide or H_2S deriv., and in general this method of prepn. of polysulfides is of greater advantage where the base forms only the polysulfide. A mixt. of I and VII was also obtained when the starting material was alc. $(\text{NH}_4)_2\text{S}_x$. In contrast to this, cold satd aq pure (approx. 5%) VI (m 143°) (100 cc. + 10 cc. addnl water) and aq $(\text{NH}_4)_2\text{S}_x$ (100 cc. prepd by satg NH_4OH (100 cc. of d 0.920) with H_2S , adding 100 cc. more of NH_4OH and heating in the absence of air with 70 g. of S) let stand several days in a sealed vessel, the ppt. (low yield) removed, pulverized in a mortar, washed successively with water, EtOH , CS_2 , EtOH , and Et_2O , and dried *in vacuo*, yields phenylbiguanide hexasulfide, $(\text{C}_6\text{H}_5\text{N})_2\text{H}_2\text{S}_6$ (VIII), orange-yellow, m 130-5°, stable when dry (did not lose H_2S when kept all summer at 26-30°), decomps very slowly in cold water and rapidly in hot water (liberation of S), and decomps slowly in EtOH (liberation of H_2S). Under no conditions was the disulfide of Romani isolated. Thus with $(\text{NH}_4)_2\text{S}_x$ contg less S (prepd as above but with 30 instead of 70 g. of S per 200 cc. of NH_4OH) only I was obtained. On increasing the quantity of S to 45 g., a mixt. of I and VII was obtained. To obtain VIII, it is necessary to adhere strictly to the proportions noted above, otherwise mixts. are obtained. These mixts. with HCl evolve H_2S and II persulfide seps., whereas pure VIII and HCl give only II persulfide. This difference is strictly analogous to the behavior of strychnine hexasulfide (IX) with acids. Besides the orange-red IX obtained by Hofmann (cf. Ber 1, 1868, 10, 1087) from cold satd alc. strychnine (X) and alc. $(\text{NH}_4)_2\text{S}_x$, the IX prepd by Schmidt (cf. Ann 180, 237) from alc. X and H_2S was also obtained, and in this latter case the compn. of the IX is, contrary to Schmidt, identical with that obtained by the method of Hofmann. Cold satd (8%) alc. $(\text{PhIN})_2\text{C NH}$ (4 g.) and S (3 g.) in suspension treated with H_2S let stand in a sealed vessel, filtered the residue washed successively with EtOH , CS_2 , EtOH and Et_2O , and dried *in vacuo*, yields diphenylguanidine hexasulfide $(\text{C}_6\text{H}_5\text{N})_2\text{H}_2\text{S}_6$ (XI), m 100-10° (to a red liquid), stable when kept dry, decomps slightly in cold water and more so in EtOH . When the S in the prepn. is diminished even as low as 0.5 g. of S (loc. cit.), XI is still obtained but the yields are smaller. XI is also obtained from aq. $(\text{PhIN})_2\text{C NH}$ H_2SO_4 or other salts and aq. $(\text{NH}_4)_2\text{S}_x$ (from NH_4OH of d 0.92). Unlike VIII, XI liberates H_2S and S when treated with acids. Therefore either it first forms a persulfide, and this is unstable, or its structure is like that of strychnine polyiodide (the more probable hypothesis). Prepd like XI, though the passage of H_2S must be much longer, di-*o*-tolylguanidine heptasulfide $(\text{C}_6\text{H}_4\text{N})_2\text{H}_2\text{S}_7$ (XII), is stable when kept dry, decomps

(VII) (cf. Meyer, C 4 10, 194) An intimate mixt. of V and excess solid KOH heated over a flame, the vapors collected as an oil, allowed to solidify, and recrystd. from C_6H_6 yields $p\text{-C}_6\text{H}_4(\text{NH}_2)_2$. By the same procedure as with 1, 3, $\alpha\text{-C}_6\text{H}_3(\text{NO})\text{OH}$ (VIII) and III yield 4,4'-formamidinoazobenzene (cf. P. C. A 17, 1632), m. 196-7°. When evapd., the C_6H_6 mother liquor yields ferric p -aminoazobenzene. In dissolving the solid product from the reaction of VIII and III, a small quantity of an unidentified brown black compd., insol. in all solvents, remained undissolved. Further expts. have been started on the reaction of isonitriles with true NO derivs., e. g., PhNO. I and PhNO boiled in C_6H_6 form besides resinous products, $(\text{PhNH})_2\text{CO}$ and a red compd., m. 212-3° which analysis indicated as an integral condensation product of 3 mols. of I, with no PhNO in combination. It is therefore not a trimer, but a true condensation compd. With PhNO III reacts like I. This investigation is to be published in more detail in a forthcoming paper.

C. C. DAVIS

Derivatives of m -xyleneols. II. Intermediate products from m -2-xyleneol (2-hydroxy-1,3-dimethylbenzene). F. M. ROWE, S. H. BANNISTER AND R. C. STORRY, *J. Soc. Chem. Ind.* 50, 79-80T (1931) cf. C. A. 25, 930—2,6,4- $\text{Me}_3(\text{NH}_2)\text{C}_6\text{H}_2\text{OH}$ (I) was prepd. in 92% yield from 2,6-Me- $\text{C}_6\text{H}_2\text{OH}$ (II) by coupling with PhNHCl and reducing with $\text{Na}_2\text{S}_2\text{O}_4$ and in 79% yield by reducing the corresponding NO compd. with $\text{Na}_2\text{S}_2\text{O}_4$. I. leaflets from H₂O or C_6H_6 , m. 137-8° (decomps.), gives a di Ac deriv., m. 160°. With 2,4-(NO_2) $_2\text{C}_6\text{H}_3\text{Cl}$ in the presence of Cu, I gave 2,4-dinitro-4'-hydroxy-3',5'-dimethyldiphenylamine, m. 193°. 2,6-Me- $\text{C}_6\text{H}_2\text{OH}$ (III), b. 182-3°, was prepd. from II and Me- SO_3H . III was best nitrated by gradually adding 103 cc. of 70% HNO_3 to 50 g. III in 75 cc. of AcOH and warming at 40-70°. This gave 2,6,4-Me $_3(\text{NO}_2)\text{C}_6\text{H}_2\text{OH}$ (IV) in 50% yield, m. 92°, which was hydrolyzed with 60% H_2SO_4 to give 2,6,4-Me $_3(\text{NO}_2)\text{C}_6\text{H}_2\text{OH}$, m. 169-70°. Reduction of IV to 2,6,4-Me $_3(\text{NH}_2)\text{C}_6\text{H}_2\text{OH}$ (V), m. 60°, was best done with Fe. SnCl_4 gave hydrolysis. V gave an Ac deriv. (VI) m. 136-7°. IV and $\text{Na}_2\text{S}_2\text{O}_4$ gave 80% of 2,6,4-Me $_3(\text{NH}_2)\text{C}_6\text{H}_2(\text{OMe})\text{SO}_3\text{Na}$ (3), from which the SO_3H group can be easily hydrolyzed to give V. V reacts with β -hydroxynaphthoic acid in the presence of PCl_5 to give 99% of the 3,5-dimethyl-4-methoxynaphthamide needles from EtOH. This gives red colors on coupling with diazo compds. VI is easily nitrated in AcOH, giving 95.5% of 2-nitro-3,5-dimethyl-4-methoxynaphthamide, pale yellow, m. 128°. The Ac group can be hydrolyzed off with NaOH or HCl, producing 1-methoxy-2,6-dimethyl-3-nitro-4-aminobenzene, m. 47°, which is converted to the corresponding 3,4-isoxazole oxide, m. 96°, by boiling with NaOCl .

G. R. YOUNG

Preparation of salicyl chloride. A. KIRPAL, *Ber.* 63B, 3190 (1930)—By adding a little AlCl_3 , $o\text{-HOC}_6\text{H}_4\text{CO}_2\text{H}$ can be made to react with SOCl_2 at a temp. at which secondary reactions do not occur when 10 g. of the acid, 7 cc. SOCl_2 and 0.02 g. AlCl_3 are kept at 45-50° the reaction is complete in 1.5 hrs. if the acid was sufficiently finely divided. Removal of the excess of SOCl_2 *in vacuo* leaves a mobile liquid solidifying in ice and showing, without further purification, the m. p. 18° given by Kopetschni and Karczag for $o\text{-HOC}_6\text{H}_4\text{COCl}$.

C. A. R.

Constituents of the volatile oil from the leaf of *Chamaecyparis obtusa*, Sieb. et Zucc., F. Formosana, Hayata or Arisan-Hinoki. I. KINZO KAFUKU, TATSUO NOZOE AND CHUTA HATA, *Bull. Chem. Soc. Japan* 6, 40-51 (1931)—Steam distn. of leaves of *Chamaecyparis obtusa* gave 0.3% of oil, slightly yellow changing to red on standing d_4^{20} 0.8988, n_D^{25} 1.4578, $[\alpha]_D^{25}$ -5.83°, acid value 0.96, ester value 12.63, ester value after acetylation 54.61. From 5 kg. of oil, 12 g. of acids and 15 g. of phenols were obtained by extrn. with 5% KOH and subsequent regeneration of the ext. in the usual manner with CO_2 and H_2SO_4 . On vacuum fractionation of the acids the main portion was an unsatd. acid apparently of the aliphatic terpene group, $\text{C}_{15}\text{H}_{26}\text{O}_2$, b. 133-5°, d_4^{20} 0.9344, n_D^{25} 1.4732, $[\alpha]_D^{25}$ 0°, mol. refraction 49.40. The remainder of the acid fraction consisted apparently of *hinokic acid*, m. 165-6°, lower fatty acids and traces of phenols. The oil remaining after extrn. of the acids on vacuum fractionation yielded 34.1% of terpenes and 11.4% of terpene alcs. Of the terpenes 50% was *d*-*isobornene*, $[\alpha]_D^{17}$ 80.4°, and 20-30% *chamene* (I), a new terpene, l., 86-8°, b. 168-70°, d_4^{25} 0.8228, n_D^{25} 1.4680, $[\alpha]_D^{25}$ 35°, mol. refraction 46.00. Satn. of the cold ether soln. of I with dry HCl gas gave the hydrochloride, b. 80-90°, d_4^{25} 1.0106, n_D^{25} 1.4783, mol. refraction 58.09. Shaking 20 g. of I with 20 cc. H_2SO_4 (1:3) at room temp. gave *isochamene* (II), b. 88-90°, d_4^{25} 0.8222, n_D^{25} 1.4726, $[\alpha]_D^{25}$ -0.27°, mol. refraction 46.42. and, contrary to chamene, fairly stable to atm. oxidation. II with alc. H_2SO_4 (1:1) polymerized to *disochamene*, b. 155-6°, d_4^{25} 0.9150, n_D^{25} 1.5134, $[\alpha]_D^{25}$ -0.7°, mol. re-

fraction 89.40. II with HNO_3 in benzene gave α -terpinene nitrosite, apparently by rearrangement of isochamene nitrosite. I gave a deep red color in Ac_2O on addn. of 1 drop of H_2SO_4 . Other terpenes found were α -pinene, α -thuyene, p -cymene, α -terpinene, γ -terpinene and dipentene. The terpene alcs. consisted of δ -l-terpinol-4, borneol, l-linalool, and its esters, and an octenol, $\text{C}_{11}\text{H}_{18}\text{O}$, n_D^{20} 0.8154, n_D^{25} 1.4441, $[\alpha]_D^{20}$ -10.62°, and its esters.

R C ELDERFIELD

Bitolyl series I A ANGELITTI WITH A BRAMILLA *Gazz. chim. ital.* 60, 967-74 (1930).—The growing interest and importance of the Ph_2 group induced A. to start a systematic series of expts on bitolyl compds. Of these, the 2,2'- and 6,6'-dimethylbiphenyl derivs. are of special importance because of the cases of optical stereoisomerism without asym atoms (cf. Mascarelli, *C. A.* 22, 68, 21, 2962, 3694, *Atti III congresso naz. chim. pura applicata* Firenze 249-62 (1930), Meisenheimer and Höring, *C. A.* 21, 2892). 4,5- $\text{Cl}_2(\text{O}_2\text{N})\text{C}_{10}\text{H}_6\text{Me}$ was converted to o,o' -dinitrobitolyl by the reaction of Ullmann and Bielecki (cf. Mascarelli and Pirani, *Gazz. chim. ital.* 41, 1, 70 (1911)) and this was partially reduced with alc. (NH_4) S by the method of Mascarelli and Gatti (*C. A.* 24, 2121) to 2-amino-2'-nitro-4,4'-dimethylbiphenyl (I) (5 g) in 7% H_2SO_4 (200 cc) diazotized at 4° with aq. NaNO_2 (16 g in 25 cc), poured slowly into a large excess of KI in 10% H_2SO_4 at 100° (cf. Gelmo, *Ber.* 39, 4175 (1906)), decolorized with SO_2 , the semi-olily product dried *in vacuo*, extd with Et_2O (a small quantity of insol. black residue is probably bitolylenciodonium iodide (II), for treated with moist Ag_2O at 60°, filtered, KI added to the filtrate, and the yellow powder recrystd from boiling water, it yields a yellowish compd which turns brown at 198° and m. 204-5° (decompn.)), while Mascarelli and Pirani (*loc. cit.*) give a m. p. of 206° for II, the ext. extd with aq. NaOH , the residue evapd, steam-distd, the 1st portion of the distillate (a small yield of 4,4'-dimethylbiphenylene oxide (III) (cf. *Ber.* 34, 3330 (1901)) the later portion (oil which solidifies to a yellow mass) dissolved in petr. ether (b. p. 70-80°), coned, and the cryst. ppt. recrystd from 80% EtOH yields 0.70 g. of 2-iodo-2'-nitro-4,4'-dimethylbiphenyl (o -iodo- o' -nitrobitolyl), pale yellow, m. 83-4°. The same procedure was used in an attempt to prep. the corresponding Br deriv., but only resinous products which could not be purified were formed. A very small yield was obtained by diazotizing I-HBr in presence of CuBr , following, in general, the method of Dobbie, Fox and Hoffmeister (cf. *C. A.* 5, 3048, 6, 593) for the prepn. of 2,2'-dibromobiphenyl. 1 (12 g) and hot water (100 cc) added to aq. HBr (45 cc. of d. 1.49) and freshly prepd. CuBr (16 g), all air removed from the flask by CO_2 , heated on a water bath, diazotized with aq. NaNO_2 (4 g in 50 cc.), (maintained at approx. 100°), heated 1 more hr. at 100°, cooled, dild. with water, the product (a black pitch contg. a yellow solid), washed with water, dried *in vacuo*, extd with Et_2O (the black residue contg. Cu will be studied later), the ext. washed with aq. NaOH , dried, evapd, the semi-olily residue (5 g) steam distd., yields 2 portions: (1) a white substance (2 g) of agreeable odor and (2) brown pitch. Portion (1) recrystd from 95% EtOH (very sol.) and then from 80% EtOH , yields 0.40 g. of III, which in this case is difficult to purify because of contamination with the chief product. The 95% alc. mother liquor from the 1st crystn. of III contains a trace of 4,4'-dimethylcarbazole. Portion (2) steam-distd. again yields an oil which partially solidifies, and which crystd. from 70% EtOH yields after long standing 0.25 g. of 2-bromo-2'-nitro-4,4'-dimethylbiphenyl (o -bromo- o' -nitrobitolyl) (IV), pale yellow, m. 73-4°. The most convenient method for prep. the Cl deriv., though the yield is very low, is the following: I (12 g) in a mixt. of coned. HCl (50 cc) and water (200 cc), diazotized hot with aq. NaNO_2 (4 g. in 50 cc.) and fresh CuCl (16 g), the black resin extd. with Et_2O , the ext. washed with aq. NaOH , dried, evapd., the oily residue (approx. 4 g) steam distd. (III first distils in small quantity as before), and the later oily distillate crystd. repeatedly from 70% EtOH , yields 0.22 g. of 2-chloro-2'-nitro-4,4'-dimethylbiphenyl (o -chloro- o' -nitrobitolyl), light yellowish, m. 83-6°. Attempts were made to prep. 2-hydroxy-2'-nitro-4,4'-dimethylbiphenyl (V), but the product could not be purified. In one case the soln. of diazonitron deriv., decompd. by bubbling in steam in the presence of urea, yielded a small quantity of III, while the chief product (also a very small yield of acidic nature) repeatedly dissolved in aq. NaOH and reprecipd. with HCl , yielded a brown-red, amorphous product which could not be steam-distd., sublimed or otherwise purified. Likewise no success was had in decomp. the diazo deriv. by pouring the soln. in hot dil. H_2SO_4 . Finally attempts were made to acetylate the reaction product, but starting from 3 g. of product (from 20 g. of I) there was obtained, after acetylation, by crystn. from 80% EtOH , a red yellow cryst. compd., m. 99-100°, mixed with a yellowish powder. The yield was so low that it could not be identified. If II is formed from small quantities of diamino deriv. as an impurity, the same explanation does not apply to the formation of

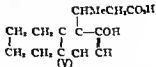
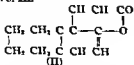
III, for the yield of III in the prepn of IV is too great to be accounted for thus. In the decompn of the diazonitro deriv., the hydroxynitro deriv., 4,6-Me(110)C₆H₂-C₆H₄(NO₂)Me 2.4 (VI) is formed initially, and this by elimination of 1 mol of HNO₂ passes to III. This hypothesis is supported by the fact that it was impossible to obtain V in the attempts in which III was formed. Furthermore it is possible that, by the action of water the NO₂ group in VI is replaced by OH, a reaction which occurs in polynitro derivs (cf Claessen, *Ger pat* 289 449). The di HfO deriv. thus formed passes readily by elimination of water, to III (cf *Ber* 25, 2745(1892)). The tendency of the Ph₂ derivs to form heterocyclic, pentatomic nuclei is observed in numerous other cases (cf Mascarelli, *C A* 3, 791, 23, 3694, Treche, *Thesis Bologna* 1911; Dobbie, Fox and Hoffmeister, *loc cit*; Mascarelli and Gatti, *loc cit*). C. C. DAVIS

Remark on the work on "aromatic thio ketones." F. BERGMANN, *Ber* 63B, 7225(1930).—In connection with the statement in the paper abstracted in *C. A.* 25, 1239 that aromatic thio ketones react with Na apparently differently from the O analogs, Schonberg has drawn B's attention to the fact that he (S) first observed the reactivity of thio ketones toward Na (*C A* 20, 364). C. A. R.

Stereochemical structure II. Optically active α - and β -methylhydrobenzoin. ROBERT ROGGE, *Biochem Z* 210, 320-9(1931), cf *C A* 24, 3000—d(+)-PhAc-CHOH (I) was prepd by Neuberg's carbolygase method (*C A* 17, 1494) and also from d(-) mandelamide and MeMgI. I and PhMgBr give d(+)-methylhydrobenzoin (α -form) (II), m 92-3° [α]_D 31° in acetone, 40° in EtOH. d(+)-Methylhydrobenzoin (β -form) (III), m 80-1° [α]_D 31° in acetone, 21° in EtOH, prepd from d(-)-benzoin and MeMgI, is not identical with II as shown by mixed m. ps. The compd. Neuberg prepd from I and PhMgBr, which he called III is not identical with II or III, but may be a phys. modification of II. d(+)-2-Methyl-2-hydroxy-1,1-diphenyl-ethanol (IV), m 92-3° [α]_D 133° in acetone, was prepd from PhMgBr and Et d(+)-lactate. IV is not identical with II or III as shown by the color with concd. H₂SO₄ and mixed m. ps. Hence I reacts as such and does not rearrange to BzMeC(OH)Ph. D. S. SEARLE

Synthesis of munjisthin. P. C. MITTER AND HAROGOPAL BISWAS, *Nature* 126, 761(1930).—Munjisthin is a dihydroxanthraquinonecarboxylic acid occurring in *Rhizoma munjistha*. It has been synthesized from 2,6-Cl(MeO)C₁₀H₆Me and C₆H₄(CO)₂H. A. L. HENCK

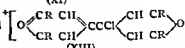
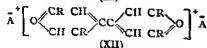
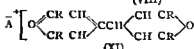
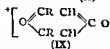
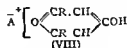
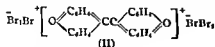
Partially hydrogenated lactones of naphthalenes. GEORG HUFMANN, *Schweiz. Anz. Z. 71*, 137-9(1931).—In the present study it is shown that relatively good yields of partially hydrogenated lactones are obtained by condensation of α - β tetrolol (I) with HO acids or β ketone acid esters in the presence of concd. H₂SO₄. Thus, malic acid and I give CO + 3 H₂O + tetrahydronaphtho- α -pyrone (II), m 131°. Similarly, AcCH₂CO₂Et and I give γ -methyltetrahydronaphtho- α -pyrone (III), m 154°. On heating with alkalis the lactone ring is opened with formation of the corresponding HO acid, stable only in the form of its salts which are obtained by evapn. of an alk. soln. of the acid. Na Hg converts the acid into the hydrogenated product (V), m. 107°, which with concd. H₂SO₄ gives the lactone. From citric acid and I there resulted γ -carboxymethyltetrahydronaphtho- α -pyrone m 189°; heating causes loss of CO₂ with formation of III.



W. O. EMERY

Conjugated double bonds. XV. Constitution of indolenine yellow. RICHARD KUHN, ALFRED WINTERSTEIN AND GEORG BALSER, *Ber* 63B, 3176-84(1931); cf *C A* 25, 1513.—As a rule, replacement of methine groups, =CH—, in a polyene chain by =N— has practically no influence on the color. Deviations from this rule occur in the initial members of homologous series and seem always to lie in the direction of stronger absorption by the N-contg. compds. An apparent exception was afforded by the violet indolenine red (I) (pure red in soln.), obtained from the methylene base (II) and salts or esters of HCO₂H, and the greenish yellow compd (III) (called indolenine yellow in the present paper), obtained according to *Ger pat.* 459,616 from II and HNO₂ and which, by analogy, was assigned the structure IV. It has been found, however, that in the condensation of bases of type II with HNO₂ in the presence of Ae₂O not only H₂O but also 1 mol HCN is split off. The yellow dye is therefore given

carbenium group they form a satd positive oxonium group and an aromatic ring, the resulting salts (VIII) therefore being colorless and much more stable than ordinary ketone salts. A and L assume that such inner satn in the direction of the betaine form (IX) has already begun in the free γ pyrones which therefore do not behave like unsatd ketones. In VIII and IX the electron seeking γ -O atom has attracted an electron from the C, as manifested by the replacement of the "double bond" in VII by the single bond in VIII and IX. The excess negative charge which the O atom thereby acquires is compensated for in VIII by the newly introduced H nucleus but is still present in IX. A diperylene system $O(CR \cdot CH)_2C \cdot C(CH \cdot CR)_2O$ (X), however, can of itself not undergo such an inner satn since there is no electron-seeking atom on the central or γ O atoms, and such a system (insoluble as it is not condensed with C_6H_5 rings as in I) is a true fulvene system and intensely colored. Its latent tendency to pyrylium formation can be brought out only by (1) introduction of an electron-seeking atom or group on the one γ C atom or (2) complete removal of 1 of the 2 pairs of electrons (i. e., of 2 electrons) of the double bond. The simplest example of (1) is where the newly introduced atom is the H nucleus of a strong acid, with formation of the *meri*-dipyrlyum salt (XI). (2) on the other hand is equiv. to an oxidation and from X or from XI, in the absence of the necessary anions, there results the *holo*-dipyrlyum salt (XII). The XI are more deeply colored than the X themselves and the XII are in general considerably less deeply colored. The XII are most simply prepd from the X with Cl or Br, the halogen acting as an oxidizing agent and also furnishing the anion (the Br anion adds another mol of Br_2 to give II). The resulting dichlorides exist in 2 forms (red and black) which rearrange into each other with extreme ease in a non-controllable manner. It is assumed that the red forms are the *holo* salts while the black forms have the structure XIII, the Cl_2 splitting into Cl^- and Cl^+ and the latter adding at 1 of the central C atoms. I, faintly greenish yellow, m 375° , dissolves in hot σ C_6H_5Cl , with blue green color (light yellow in the cold), forms a blue green melt, shows intense light blue fluorescence under the quartz lamp. Thioxanthone, very faintly yellow, m 213° , shows yellow fluorescence in ultra violet light. Thioxanthione, dark needles with green surface luster, m 176° , sol in C_6H_6 with red color (by transmitted light, green by incident light or in thin layers), in $CHCl_3$ with pure green, in concd H_2SO_4 with dark orange color. IV, m 360° , shows blue fluorescence in ultra-violet light. Thioxanthone gives on reduction according to F. Mayer a compd m 346° but behaving toward Br like IV. Dithioxanthylene disulfone, from IV and H_2O_2 in AcOH, gradually darkens between 380° and 500° , melts at red heat on Pt, does not react with Br soln or vapor even on heating. Dixanthionium perbromide crimson, darkens 220° , decomp 245° , regenerates IV with Zn dust and AcOH. *meri*-Perchlorate from V, obtained by boiling finely divided V (pptd from C_6H_5N with water) in AcOH under CO_2 , filtering and adding 70% $HClO_4$, brown black, decomp $287-8^\circ$ with considerable violence. *holo* Perchlorate, from V and $HClO_4$ at about 200° , red, deflagrates vigorously when heated. *meri* Perchlorate from VI, red needles, changes in hot concd H_2SO_4 or 70% $HClO_4$ or on treatment in cold AcOH with $HClO_4$ and Br from deep red to green. Tetraphenyldithiopyrlylene (XIV) in $CHCl_3$ adds only 2 atoms Cl in CCl_4 , forming 2 mutually interconvertible dichlorides (red and black), the red form darkens around 80° , both sinter 100° and decomp $300-10^\circ$. Dichlorides of V; the red rapidly changes into the black form which sinters around 100° , m 305° . With Br in $CHCl_3$ XIV gives a hexabromide, golden yellow, loses Br at $60-100^\circ$ on slow heating, becomes smeary at 120° .



Some 2,4-substituted derivatives of pyridine C A R
RODERICH GRAF (WITH ERNST

LEDERER-PONZER AND LEOPOLD FREIBERG) *Ber* 64B, 21-6(1931); cf. *C. A.* 23, 837—4-Chloro- (I) and 4-iodopicolinic acid (II) were degraded by the Curtius method through the hydrazide and azide to 4-chloro- (III) and 4-iodo-2-aminopyridine (IV). Attempts to convert the I by the Naegeli method through the chloride directly into III failed because of the slight reactivity of the chloride with NaN_3 . In C_6H_5 the chloride with N_3H_4 , H_2O gave di-4-chloropicolinic hydrazide (V). The simple hydrazide (VI) was best prepd by converting the I with SOCl_2 into the chloride, then with excess of MeOH into the Me ester which gave the VI almost quant. Attempts to degrade the VI directly to the urethan, without isolating the azide, by boiling in abs alc. with AmONO gave a difficultly sol., unreactive substance which carbonized at 272° without melting. On the other hand, the azide boiled with dil AcOH gave III directly. By the Schotten-Baumann method, III gives the di-Bz deriv., converted by long boiling with alc into the mono-Bz deriv. I, faintly yellowish, m $180-1^\circ$, was obtained in 50-55% yield by refluxing anhyd picolinic acid HCl in SOCl_2 for 10 days and decomp., the resulting chloropicolinyl chloride HCl with H_2O . V, faintly yellowish, m $269-71^\circ$ (decompn). VI, from the chloride (m 46°), m $167-8^\circ$ (18 g from 185 g of I), benzal deriv., m 178° . Azide, from VI m cold N HCl with aq KNO_3 , m 92° , deflagrates on rapid heating, gives with I AcOH on the water bath N_2CO_2 and III (3.2 g from 7.4 g VI), m $130-1^\circ$, easily volatile with steam, picrate, yellow, m $243-4^\circ$. *N*-Carb-ethoxy deriv. of III, from the azide and boiling abs alc., m 161° , best hydrolyzed by const. boiling III, which gives III. III, m $200-7^\circ$, soon turns yellow in the air. *N,N*-Di-Bz deriv. m $165-6^\circ$. *N*-Bz deriv., m $120-1^\circ$. *N*-Ac deriv., from III and boiling Ac_2O , m $115-6^\circ$. *2-HO* compd., from III in dil H_2SO_4 with the calcd amount of KNO_3 , m 184° , only slightly sol in dil acids. 2,4-Dichloropyridine, from III by the Sandmeyer method (about 1 g from 5 g III), b about 184° , m 0° , almost insol in dil acids, exceedingly volatile with steam, forms a cryst double salt with HgCl_2 but no picrate. II from I refluxed in III (d 1.7) with some red P, m 160° (decompn), cannot be converted into the chloride with SOCl_2 , because of exchange of the I for Cl. The Me ester, from II and $\text{MeOH-H}_2\text{SO}_4$, yielded with concd alc NH_3 the amide, m 158° , and with alc N_3H_4 , H_2O the hydrazide, m $160-1^\circ$ (benzal deriv., m $207-8^\circ$). IV (through the amide, m 89° (foaming)), m $163-4^\circ$, picrate, yellow, m $253-4^\circ$; *N*-carbethoxy deriv., m 167° . *N,N*-di-Bz deriv., m $176-7^\circ$, *N*-Bz deriv., m $167-8^\circ$; *N*-Ac deriv., m 150° . 2-HO compd., m 195° . C. A. R.

Imide and amide chlorides of non-aromatic acids. VI. A new way in the quinoline series. JULIUS V. BRAUN AND ALBRICHT FIEBENS (WITH L. SCHNITZSPAHN) *Ber* 63B, 3191-203(1930), cf. *C. A.* 24, 2999—Wallach assigned the structure $\text{MeC}(\text{NR})\text{CH}_2\text{CCl}$ NR to the compds obtained by condensation, with loss of 1 HCl , of 2 mols of the imide chlorides, MeCCl NR, derived from AcOH , but it was later shown that they are amidines, $\text{MeC}(\text{NR})\text{NRCCl}\cdot\text{CH}_2$, *s. c.*, that the condensation takes place between a C and N atom. It was thought that imide chlorides of the type $\text{R}'\text{CH}_2\text{CCl}$ NR undergo a similar condensation, since Wallach, and later Bischoff and Walden, had found that $\text{ClCH}_2\text{CONHPh}$ gives a compd $\text{C}_{12}\text{H}_{12}\text{N}_2\text{Cl}_2$ (I), whose formation through an intermediate imide chloride, ClCH_2CCl NPh, could readily be explained by the equation $2 \text{C}_6\text{H}_5\text{NCl}_2 = \text{I} + \text{HCl}$. I, however, is characterized by an unexpectedly strong, bright yellow color and by the surprising ease (treatment with cold Na_2CO_3) with which it loses another mol of HCl , and in addn to the Wallach formula, $\text{ClCH}_2\text{C}(\text{NPh})\text{CHClCCl}$ NPh, Bischoff and Walden suggested the structure of a HCl salt $\text{ClCH}_2\text{C}(\text{NPh})\text{CHCl}$ CII NPh HCl . Recently, in connection with

the prepn of imide chlorides from chloroacetanilides with substituents in the *o*-position to the N (*C. A.* 24, 2999), similarly bright yellow compds were obtained, and although analysis showed they were formed by loss of 1 mol HCl from 2 mols imide chloride the difference in color between the AcOH and $\text{ClCH}_2\text{CO}_2\text{H}$ derivs. was too great to be explainable on the basis of the amidine structure for both. A more thorough study has now shown that I is indeed a HCl salt of a colorless base $\text{C}_{12}\text{H}_{12}\text{N}_2\text{Cl}_2$ (II) which when heated with mineral acids under pressure splits off only 1 mol PhNH_2 ; on less-energetic treatment with concd HCl it yields a weakly basic substance $\text{C}_{10}\text{H}_8\text{NCl}_2$ (III) and on more vigorous treatment a phenolic compd $\text{C}_8\text{H}_6(\text{OH})\text{NCl}_2$ (IV) which regenerates III with PCl_5 . III with H and P loses all its Cl and gives pure quinaldine. The imide chlorides derived from the chloroacetylides therefore condense with 2-fold elimination of HCl to give, not amidines like the acetylides, but quinolines. It has thus far been impossible to force the PCl_5 reaction with $\text{ClCH}_2\text{CONHPh}$ in the amidine direction and that with AcNHPh in the quinoline direction. Derivs. of EtCO_2H and $\text{PhCH}_2\text{CO}_2\text{H}$ also yield quinolines. The new reaction should prove to be one of the

most fruitful for the prepn of quinolines, for the ring formation takes place very readily and in part quantitatively and the products can undergo numberless further transformations. 2-Chloromethyl-3-chloro-4-anilinoquinoline-HCl (II), decomps 222° (B and W give 212°), is obtained in 90-5% yield from $\text{C}_6\text{H}_5\text{CONHPh}$, cooled in running H_2O treated with 1 mol PCl_5 allowed to stand 12 hrs. and washed with H_2O , the free base (II), m 131°, is obtained by taking I up in a little cold $\text{C}_6\text{H}_5\text{N}$ and quickly pptg the II with dil NH_4OH or Na_2CO_3 . The course of the reaction of II with aq HCl under pressure depends on the length of heating, the temp and the amt of HCl , for not alone is the NHPh group replaced by Cl and the latter by H_2O but the Cl in the $\text{C}_6\text{H}_4\text{Cl}$ side chain can independently be replaced by H_2O . The 2-chloromethyl-3-chloro-4-hydroxyquinoline (IV), m 303°, is obtained in about 80% yield from II and 4 parts concd HCl heated 12 hrs at 160-5° and 1 hr at 175-80°; it dissolves only in excess of concd acid and is reprecipitated by diln. with H_2O . Along with the IV is formed the alkalinsol III, the amt of which is the greater the shorter the length of heating (46% each of III and IV after 14 min. at 170-50°). With only 1.5 parts by wt. of concd HCl at 165°, II gives 77% IV and the H_2O sol 2-hydroxymethyl-3-chloro-4-hydroxyquinoline, m 283°, readily obtained by boiling IV several hrs with aq alkali. The 2-chloromethyl-3,4-dichloroquinoline (III), m 119-20°, is readily obtained by heating IV a few min at 130-40° with somewhat more than 1 mol PCl_5 and a little POCl_3 . The 4-Cl and still more the Cl atom in the side chain are reactive and therefore heating III in aq alc a few hrs with 1 mol alkali gives the alkalinsol, 2-hydroxymethyl-3,4-dichloroquinoline, m 44°, b₁₁ 150-4°. The compd obtained by Buschoff and Walden from I with Zn dust and AcOH , which they believed to have the compn $\text{C}_{11}\text{H}_9\text{N}_2\text{Cl}$, is in reality 3-chloro-4-anilinoquinoline, it is best obtained from II with excess of Sn and concd HCl on the H_2O bath and m 172°, picrate, yellow, m 229°. 3-Chloro-4-hydroxyquinoline, from IV in much boiling AcOH with 10 parts Zn dust, m 340°, converted by PCl_5 POCl_3 into 3,4-dichloroquinoline, m 322°. With 30 atoms Na and alc. II yields py tetrahydroquinoline, b₁₁ 125°. Short warming of II with aq alc. KOH gives 2-hydroxymethyl-3-chloro-4-anilinoquinoline, m 93-4°, with KCN in aq alc. is obtained the 2-cyanomethyl compd, m 156°, which readily condenses with H_2N in alc. and a few drops of alkali to form a compd $\text{C}_{11}\text{H}_{10}\text{N}_2\text{Cl}$, yellow, m 192°. With PhNH_2 , II gives 2-anilinomethyl-3-chloro-4-anilinoquinoline, m 192°, isolated as the yellow HCl salt, and acetylated on the new N atom by hot Ac_2O , giving an Ac deriv m 200°. With 3 mols NHMe_3 in alc. II gives 2-dimethylaminomethyl-3-chloro-4-anilinoquinoline, b₁₁ 215-20°, m 93°, picrate, m 190-3°, with 21 atoms Na and alc. the base gives PhNH_2 and 2-dimethylaminomethyl py tetrahydroquinoline, yellowish, b₁₁ 148-53°. 2-Dimethylaminomethyl-3,4-dichloroquinoline, similarly obtained from III, m 62°, picrate, red yellow, m 177°. 2-Piperidinomethyl-3-chloro-4-hydroxyquinoline, from IV, m 233-5°, sol in both acids and alkalies. II and III also react with tertiary bases through the reactive ClH_2Cl group. II with $\text{C}_6\text{H}_5\text{N}$ gives the quaternary compd $\text{C}_{11}\text{H}_{10}\text{N}_2\text{Cl}$, deep yellow, m 245°, isoquinoline compd, $\text{C}_{11}\text{H}_{10}\text{N}_2\text{Cl}_2$, m 258°. Both give with aq alkali slimy ppts of the corresponding pseudo bases which, when fresh, redissolve in HCl to form the original salts but undergo change on standing. The yellow compd $\text{C}_{11}\text{H}_{10}\text{N}_2\text{Cl}$, m 272°, obtained by Buschoff and Walden from p-MeC₆H₄NHCOCH₂Cl and PCl_5 is readily obtained from p-MeC₆H₄NHCOCH₂Cl and PCl_5 is readily obtained from p-MeC₆H₄NHCOCH₂Cl N-(Chloroacetyl)-p-toluidine, m 193°, likewise gives with PCl_5 a yellow magma which, pptd from cold $\text{C}_6\text{H}_5\text{N}$ with dil Na_2CO_3 yields the di I analog, $\text{C}_{11}\text{H}_{10}\text{N}_2\text{Cl}_2$, of II, m 184° (HCl salt, yellow, m 232°; picrate, m 202°). 2-Diethylaminomethyl compd, $\text{C}_{11}\text{H}_{12}\text{N}_2\text{Cl}$, faintly yellow, m 152° (picrate, yellow, m 178°). m $\text{C}_6\text{H}_5\text{N}$, NHCOCH_2Cl gives practically only 1 product, $\text{C}_{11}\text{H}_{10}\text{N}_2\text{Cl}_2$, m 179°, reacting smoothly with p-EtOC₆H₄NH₂ to form 2-p-phenethidinomethyl-3-chloro-4-m-chloroanilino 7-chloroquinoline, m 131°. Likewise, 3,4-CiMeC₆H₃NHCOCH₂Cl gives a practically homogeneous quinoline $\text{C}_{11}\text{H}_8\text{N}_2\text{Cl}_2$, m 174°. These results with PhNH_2 and its m- and p-substituted derivs showed that the compds recently isolated in very small amounts along with the imide chlorides from o-substituted chloroacetanilides are quinolines and not amidines. These compds have the following m ps: from o-toluidine, 160°; o-anisidine, 196°; o-CiC₆H₃NH₂, 110°; o-BrC₆H₃NH₂ does not crystallize well (HCl salt, m 135°). 2-Ethyl-3-methyl-4-anilinoquinoline HCl , from EtCONHPh and PCl_5 (yield, more than 50%), yellow, m 217-20°, free base, m 178°; HCl salt, yellow, m 219-20°; picrate, m 221°. The base heated 2.5 hrs. at 155° with 10 parts concd HCl gives 1-nH₂ and 2-ethyl-3-methyl-4-chloroquinoline, m 40-3°. If the temp during the heating is raised to 180-90° there is obtained the 4-HO compd, only slightly sol in alkalies but easily sol in dil mineral acids, m 284°, converted back into the 4-Cl compd by PCl_5 - POCl_3 at 140° and reduced by Na-EtOH to the py tetrahydro-2-ethyl-3-methyl-

quinoline, b_p 136–40° (HCl salt, m 193°) 2-Benzyl-3-phenyl-4-anilinoquinoline HCl (80% yield), yellow, m 209°; free base, m 172°, *picrate*, m 188°. Heated 6 hrs at 170–200° with concd. HCl the 4-Cl compd gives the 4-HO deriv, m 308° 2-Benzyl-3-phenyltetrahydroquinoline, light yellow, b_p 190–2°, non-volatile with steam C. A. R.

Azo compounds from trimethylmethyleneketopyrazine. VII. E. PRINCIVALLE. (*Ann chim ital* 60, 963–6(1930), cf. C. A. 24, 3792—2,5-Dimethyl-6-keto-1,6-dihydropyrazine couples easily with diazonium salts (cf C. A. 23, 3472), but if the formation of a lactic group is prevented by substitution of the imine H by a radical, no coupling takes place (cf C. A. 24, 3792). Azo compds are also formed by the reaction of 1,2,4-trimethyl-5-methylene-6-keto-1,6,4,5-tetrahydropyrazine (I) (cf C. A. 24, 1383) and diazonium salts, since diazonium salts react easily with compds, which contain multiple aliphatic bonds (cf Meyer, C. A. 14, 1330, Quilico, C. A. 23, 597, 3675). The reaction takes place by addn to the C CH₃ group, with subsequent elimination

of H₂O thus $\text{OC NMe CMe CH NMe C CH}_3$ (I) + PhN NOH \rightarrow —NMe.C.

(OH)CH₂N NPh \rightarrow —NMe.C CHN NPh I reacts also with PhINNH₂ (II), with formation of the same azo compd as that by the action of PhN.NCl, and the reaction takes place in a similar way. The basic products are stable compds, in general of an intense red color, and with H halides form salts; the structure is not certain (see later). I and II (equimol quantities) heated in 20% AcOH for 48 hrs at 50°, cooled, neutralized with NaOH, and the product recrystd from EtOH, yields 1,2,4-trimethyl-5-phenylazomethylene-6-keto-1,6,4,5-tetrahydropyrazine (III), red, m 201° (decompn). A better yield of III is obtained with PhN₂Cl though the procedure is more tedious, i. e. sufficient NaOH is added to 1,2,5-trimethyl-6-keto-1,6-dihydropyrazine-MeI soln to liberate the base, the mixt is cooled, PhN₂Cl (calcd quantity) is added dropwise, the product is let stand until sepn. is complete, the ppt (a mixt of III-HCl and IV) recrystd from dil HCl, compd with NaOH, and the III recrystd from EtOH, dil HCl or dil HI. III and dil HI allowed to crystallize yield the III salt C₁₁H₁₁ON₂Cl (IV), red, m 223° (decompn). By the same procedure as in the previous prepn, there is obtained from *p*-MeC₆H₄N₂Cl, after recrystn from EtOH, 1,2,4-trimethyl-5-*p*-tolylazomethylene-6-keto-1,6,4,5-tetrahydropyrazine (V), red, m 214° (decompn). V and dil. HI yield the III salt C₁₁H₁₁ON₂Cl (VI), brilliant green, m 226° (decompn). Likewise from V and dil HCl is obtained the HCl salt, C₁₁H₁₁ON₂Cl (VII), green, m. 205° (decompn). The salts IV, VI and VII may

have either or both of two structures, $\text{OC NMe.CMe.CH.NHMeI.C:CHN:NPh}$

(VIII) or $\text{OC NMe CMe:CH.NMeI:CCH NNHPh}$ (IX). VIII would be formed by the normal addn of the halogen acid, with conversion of N from trivalent to quinquivalent. IX would be formed from VIII by subsequent isomerization. It is very probable that IX represents the form of IV, VI and VII, which means that they have a structure analogous to those formed from 2,3,3-trimethylindolenine-MeI (cf Rosenbauer and Feilner, C. A. 21, 407). Further expts are in progress to settle this problem.

C. S. DAVIS

Strychnos alkaloids. LVII. Oxidation of tetrahydrostrychnine and some derivatives. HERMANN LEUCHS (IN PART WITH FRITZ KRÖHNKE). *Ber.* 63B, 3181–9(1930), cf. C. A. 25, 705—Whereas in the oxidation of hucine, the aminostrychnines, strychnine and strychnidine (I) by CrO₂ only products which had been attacked in the C₁₁H₁₁ nucleus could be isolated, the oxidation of tetrahydrostrychnine (II) proceeded differently. While here, too, there was obtained a compd. C₂₁H₂₉O₄N₃ (III), analogous to diketonucidine, isolated in 7.4% yield as the perchlorate, there was also obtained 14% of an amino acid C₂₁H₂₉O₄N₃ 4H₂O (IV), whose compn shows that the C₁₁H₁₁ nucleus is still intact. IV had already been observed in the oxidation of I, but in such small and widely varying yields that its formation was probably due to the presence of H in the crude I. The oxidation of the II consists in the taking up of two O and loss of four H atoms, since a CO₂H group is formed, there may have occurred the change —CH₂(OH) into CO₂H, with which would have to be related that of —CH₂ NH— into —CO NH— or —C(OH) N—. For in strychninic acid, with the same groups —CH₂ NH— and H₂O-C—, these are anhydriized by mineral acids. The negative semicarbazone test would indicate that the compd is not a ketone. Oxidation of a (b) N.CH₂— residue would still leave a salt-forming group in (a) NH CH= but would not explain the lactam ring remaining open. To test these views, the Ac derivs

of II were studied. The *O,N*-di Ac deriv (V) is quite stable toward CrO_3 , it consumed less than 19 equivs giving chiefly, with loss of the *O* Ac group, an amino acid $\text{C}_{11}\text{H}_{19}\text{O}_4\text{N}_2$ (VI) isolated in the perchlorate in 55% yield. From the compn of VI, the reaction must have been limited to sapon and conversion of $-\text{CH}_2(\text{OH})$ into $-\text{CO}_2\text{H}$, i. e., VI must be Δ -acetyl-strychnine acid, attempts to prep it from Ac_2O and an alk. soln of the acid gave only strychnine. The sapon precedes the oxidation, and the *N* Ac deriv likewise gives VI (in 77% yield with only 14 equivs CrO_3), 10 equivs KMnO_4 in Me_2CO at 0° gave no cryst products. Attempts to oxidize II with Br gave first an aromatic di Br deriv (*o, p* to the (a)N) (VII) which then changed into resinous substances. Perchlorate of I, needles and prisms with 1 H_2O , $[\alpha]_D^{25} -10.1^\circ/\text{d}$ (H_2O); HCl salt, $[\alpha]_D^{25} -13.5^\circ/\text{d}$. Perchlorate of II, $[\alpha]_D^{25} -74.9^\circ/\text{d}$; HCl salt, $[\alpha]_D^{25} -80.6^\circ/\text{d}$; free base needles with 2 H_2O , m 202° (vacuum), *N*-Ac deriv, prisms with 1 H_2O , m $157-9^\circ$ and anhyd, 197.9° (vacuum), perchlorate, $[\alpha]_D^{25} -101.6^\circ/\text{d}$; HCl salt, $[\alpha]_D^{25} -122.3^\circ/\text{d}$. V, tables and prisms with 1 H_2O , m around 100° and, anhyd, $135-42^\circ$ $[\alpha]_D^{25}$ in 1 mol 0.1 *N* HCl $112.2^\circ/\text{d}$. perchlorate prisms with 1.5 H_2O , $[\alpha]_D^{25} -96^\circ/\text{d}$. VII, mters 215° m 218.10° ($250-2^\circ$ in vacuo), shows no basic reactions and does not react with FeCl_3 , dissolves in CrO_3 , H_2SO_4 with red violet color, is unchanged by boiling H_2O , HBr salt needles losing 10.1% in wt at 100° and 15 mm; HCl salt, $[\alpha]_D^{25} -23.2^\circ/\text{d}$. IV, needles or leaflets with 4 H_2O (erroneously given as 3 H_2O in an earlier paper (C A 25, 705)) sol in *N* KHCO_3 but not in *N* NaOAc , does not m 320° ; HCl salt, $[\alpha]_D^{25} -21.0^\circ/\text{d}$. Perchlorate of III, tables and octahedrons with 2 H_2O , $[\alpha]_D^{25} -73.8^\circ/\text{d}$, free base, faintly yellow prisms, losing (under 15 mm) 1.5% in wt at 75° and 2.3% at 115° , mters 230° m 243.5° ($215-7^\circ$ in vacuo). Oxidation with 24 equivs CrO_3 gave more than 10% IV and less than 5% salt of III; pretreatment of the II with $\text{IcNH}_4(\text{SO}_4)$, decreased the yields. Perchlorate of IV, $[\alpha]_D^{25} -118.5^\circ/\text{d}$, (H_2O), free acid, prisms with neutral reaction, mters $180-200^\circ$, becomes foamy 220° , decomp 305° . C. A. R.

Sinomenine and disinomenine XXIII The identity of β -tetra-hydrodesoxycodeine and dihydrothebaine KAKUJI GOTO AND SHINGO MITSUI *Bull Chem Soc Japan* 6, 33-9(1931) -Kondo and Ochiai (C A 24, 3213) obtained desoxytetrahydrosinomenine, the *d* form of dihydrothebaine (I), by Clemmensen reduction of sinomenine (II) whereby the phenolic OH group is apparently removed and the CO group reduced to a secondary alc. Goto (C A 24, 1345) attempting to repeat this obtained desoxydemethoxydihydrosinomenine (III), the *d* form of β -tetrahydrodesoxycodeine (IV). The retention of the phenolic OH and reduction of the CO group to CH_2 is now confirmed by reduction of demethoxydihydrosinomenine to III by means of IcCl_3 and subsequent catalytic reduction of the ketone dichloride. By subjecting various derivs of II to the Takai electrolytic reduction G and M were able in no case to remove the phenolic OH and obtain an alc. base corresponding to the dihydrothebaine of Speyer and Siebert. The identity of I and IV is further supported by the identity of the des *N*-Me bases, m $147-8^\circ$, and the *N* free bases, m $107-8^\circ$, from both sources. It is proposed to substitute the name *dihydro-sin thebaine* for K and O's *dihydrothebaine*, designating I, III and IV. The easy reduction by the Clemmensen method of the MeO and corresponding OH groups in all derivs of I contg the CO group makes it seem probable that these groups are vicinal. R. C. LIDSEFIELD.

The saponin of ginseng bark. I. E. WEDERKIND AND W. SCHICKLE *Z physiol Chem* 195, 132-8(1931) -The crude saponin was hydrolyzed with 5% H_2SO_4 and the insol. amorphous end saponin, for which the name *guagenin* is proposed, was used in the prepn of derivs. Ac_2O converted it into a cryst. *acetylguagenin*, m 268° , $[\alpha]_D^{25}$ in CHCl_3 , 74.5° , from which the *Me ester*, m 220° , $[\alpha]_D^{25}$ in CHCl_3 , 69.2° , was obtained by treatment with CH_3N_3 . By sapon the Ac deriv with KOH in MeOH the *guagenin*, $\text{C}_{21}\text{H}_{39}\text{O}_7$, m 310° , was obtained cryst. This cryst product yielded a *Me ester*, m 190° , $[\alpha]_D^{25}$, 73.9° , and an Ac deriv identical with that prepd from the amorphous saponin. A. W. DOX.

Synthetic experiments on the constitution of the hule pigment III. Synthesis of hydroxypyromethenes, and some derivatives of coproporphyrin I. HANS FISCHER AND WALTER IRÖDIS. *Z physiol Chem* 195, 49-89(1931), cf C A 18, 2515 -A study was first made of the reactivity of the α Br m $4,3',5'$ trimethyl-3,4'-dipropionic acid 5-bromopyromethene and its *Me ester* (I), since a replacement of the Br by OH should yield a product very closely related to xanthobilirubin acid. I reacts with PhNH_2 to form $4,3',5'$ trimethyl 3,4'-di(propionic *Me ester*)-5 phenylaminopyromethene-

HBr, m 180°. The free base, m 145°, forms a complex *Cu* salt, m 120°. Treatment of I with *MeNH*, does not replace the *Br* but converts the ester groups into the corresponding amide, yielding 4,3',5'-trimethyl-3,4'-di(propionic methylamide)-5-bromopyrromethene (II), m 212° (decompn). A similar reaction occurs with *NH*, in *MeOH*, which yields 4,3',5'-trimethyl-3,4'-di(propionamide)-5-bromopyrromethene, decompn 260–70°. Treatment of II with 10% *KOH* in *MeOH* replaces the *Br* by *MeO*, forming 4,3',5'-trimethyl-3,4'-di(propionic methylamide)-5-methoxypyrromethene, m. 225°. I yields a similar *MeO* deriv, viz 4,3',5'-trimethyl-3,4'-di(propionic *Me* ester)-5-methoxypyrromethene (III), m. 88°. The *OEt* deriv representing hydrolysis of the *Br* in I, viz, 4,3',5'-trimethyl-3,4'-di(propionic *Me* ester) 5-hydroxypyrromethene, m 180°, may be prepd directly from I by heating the *AcOH* soln with *AgOAc*, or from III by hydrolysis of the *MeO*. This is accomplished by heating III with *MeONa* at 170–80° and reesterifying the carboxyls with *Me₂SO₄*, or better by resorcinol fusion of III at 170–80°. This hydroxymethene in *CHCl₃* gives a green color reaction with diazobenzenesulfonic acid and *HCl*. Heated to 170–80° with *HBr-AcOH*, 2 mols condense to coproporphyrin in small yield. The di-*Et* deriv corresponding to the free acid of I, when heated with *KOH* in *MeOH*, yields 4,3',5'-trimethyl-3,4'-diethyl-5-methoxypyrromethene, m 70°, from which 4,3',5'-trimethyl-3,4'-diethyl-5-hydroxypyrromethene, m 243°, is obtained in small yield by resorcinol fusion. The mother liquors remaining from the prepn of the *HBr* salt of I (free acid), when evapd spontaneously over a period of several months, gradually formed coproporphyrin I, which was isolated and identified as the tetra-*Et* ester, m 226°. In the prepn of coproporphyrin tetra-*Me* ester by citric acid fusion a mother liquor remained from which the *Me* ester of coproporphyrin IV, m 177°, was isolated. Porphyrin formation occurs also in small yield when the *HBr* salt of the free acid of I is treated with *H₂SO₄*. Coproporphyrin tetra-*Me* ester I and isuroporphyrin octa-*Me* ester II yield complex *Co* salts, m 270° and 316°, resp. The copro ester gives a *Ag* salt, m 286°, which may be sapon to the free acid by boiling with *KCN* in 90% *EtOH* and reesterified by shaking with *Me₂SO₄* and *Na₂CO₃*. Treatment of the *Ag* salt ester with *Fe(OAc)₃* and *AcOH* contg a little *NaCl* replaces the *Ag* by *Fe*, yielding the tetra-*Me* ester of coprohematin, m 245°. *Na₂CO₃* converts the latter into coprohematin tetra-*Me* ester, m. 215°. Reduction of this ester with *Na* and *AmOH* converts it into the corresponding chlorintetracarboxylic acid which dissolves in *Et₂O* to a stable green soln., but its alk soln soon reverts to the red porphyrin. Exposure of an aq colloidal suspension of coproporphyrin to summer sunlight for several days converts it into a chlorin which differs from the synthetic chlorin in that its alk soln is stable to atm *O*. Nitration of the copro ester with concd *HNO₃* yields a dinitrocoproporphyrin, and hematinic anhydride as a by product. This dinitro deriv forms a tetra-*Me* ester, m 191° (complex *Cu* salt, m 200°, complex *Ag* salt, m. 240°). If the esterification is performed by refluxing with 5% *HCl* in *MeOH* the product is a mononitrodihydroxyporphyrin ester, m 206° (complex *Cu* salt, m 245°). Attempts to prep an aminoporphyrin by reduction of the *NO₂* were unsuccessful. After treatment of the dinitro deriv with *Na-Hg* the products isolated were coproporphyrin, dinitroporphyrin, and mononitroporphyrin, from which a tetra-*Me* ester, m. 204°, was prepd. The mono- is more stable than the dinitro deriv, and can be esterified at refluxing temp. Finally a tetrahydrazide, was prepd from the tetra-*Me* ester of coproporphyrin and *N₂H₄*, *H₂O* in *MeOH*. Its *Fe* salt was obtained in the same manner from coprohematin ester. The tetraazide of coproporphyrin was prepd. by treatment of the hydrazide with *NaNO₂*, and *HCl* with ice cooling. It explodes at about 90°. *MeOH* converts the azide into a tetraurethane, m 190°. This is a remarkably stable substance, not attacked by concd. *H₂SO₄*, or by resorcinol fusion at 200°. It is apparently sapon by *HCl* at 200°, but the expected tetraaminoporphyrin could not be isolated in a pure state.

A. W. Dorr

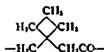
Blossom xanthophylls. The pigment of the sunflower. L. ZECHMEISTER AND P. TUZSON *Ber.* 63B, 3203–7(1930) —The polarimetric method is at present probably the most reliable means of distinguishing closely related pigments of the xanthophyll type from each other. The non homogeneity of leaf xanthophyll was proved by the isolation of analytically pure preps of widely varying rotations (137–92°) from stinging nettles; the maize pigment, zeaxanthin, can readily be distinguished from the xanthophyll of leaves by its *l* rotation, and the lutein of egg yolk by its characteristically low *d* rotation (72°). In the search for other natural xanthophylls differing in rotation from that of leaves, it has been found that the well-crystd pigment (I) of the sunflower (C. A. 25, 314) has a low *d* rotation (75°) identical, within the exptl errors, with that of the egg lutein, and as its other consts also agree with those of lutein, there can hardly be any doubt that the two substances are identical. The sunflower

seems to contain the lutein mostly in the form of a cryst. ester. The I (0.234 g from 300 g dried petals) was isolated by letting the fresh petals stand 1 day under 90-6% EtOH drying at 35°, powdering, percolating with C_6H_6 , concg in *vacuo* at 40°, adding 98% alc (which ppt a colorless substance), filtering, adding H_2O , extg with H_2O , letting the washed H_2O ext stand 20 hrs with concd KOH in MeOH, concg the H_2O layer (dried with Na_2SO_4) in *vacuo* in a current of CO_2 , pptg with petroleum ether, extg with boiling 70% EtOH and crystg the insol residue from MeOH. It m 193° (cor) $[\alpha]_D^{25} 75^\circ$ (CHCl_3) spectrum in 96% EtOH (5 mg in 1 l, 10-mm layer), 486-72, 455-5-442 μ about 420. C. A. R.

The iron salt of pheophorbide b. OTTO WARBURG. *Ber* 64B, 682-3 (1931) — A mixt of 1 g l e powder, 1 g NaCl and 60 cc glacial AcOH is boiled in an atm of CO_2 , cooled and filtered. The l e salt of pheophorbide b is obtained in 5.20 mg yield by adding 1 g pheophorbide b to the above soln, passing a current of CO_2 through it and heating for 10 min on a water bath. The crude product deposited on cooling is centrifuged and washed with 50% AcOH and H_2O . It may be crystd by dissolving 100 mg in 15 cc hot propionic acid, filtering and adding 7.5 cc hot 0.5% HCl. The crystals which sep on cooling are probably $(\text{CH}_3\text{CH}_2\text{CO})_2\text{C}_4\text{H}_7\text{N}_3\text{O}_4\text{FeCl}$. The Fe salt of pheophorbide a may be similarly prepd. RUTH BEACON

Isolation of phytosterolin from wheat embryo. NOBUZO NAKAMURA and AKI YOSHIIKAWA. *Sci Papers Inst Phys Chem Research* (Tokyo) 15, Nos 280-91, 137-41 (1931) — From the ether ext of wheat embryo 0.2% of phytosterolin was isolated. It is identical with the glucoside spuranol of phytosterol from other materials (C. A. 7, 2218). It forms needles (from AmOH), m 285-90°. Ac deriv m 168°, $[\alpha]_D^{25} -22.4^\circ$ (in CHCl_3). Bz deriv m 198°, $[\alpha]_D^{25} 17.24^\circ$ (in CHCl_3). ALFRED BURGER

Some relations between constitution and odor. A. ANGELL. *Gazz chim ital* 60, 239-46 (1930) — In view of the odorous characteristics of aromatic musks and of violet ketones, it is surprising that the ketones first synthesized by Ruzicka possess such remarkable odors, for they contain no Me groups around the CO residue. Consequently, because of the size of the ring resulting from a great no of methylene groups, deformation of the ring itself probably occurs in such a way that lateral nodes are formed. These nodes are made up of residues which function in the same way as do the Me groups in the musks and ketones mentioned. In this case there would be a twisting as represented by the structure



and this seems likely from the fact that, though x ray examn of the ketones of Ruzicka furnish no evidence, Lee (cf C. A. 21, 3189) has noticed anomalies in certain long-chain compds which are greatest with 15 C atoms. Furthermore Lee and van Rysselberghe (cf C. A. 24, 59) suggest a spiral form in long-chain compds. To throw light on this problem, the chem behavior of some cyclic ketones was studied, particularly the influence of certain substituents (alc, radicals) on the reactions which are characteristic of these ketones. If the hypothesis already suggested, i. e., that in proximity to the CO group the methylene chain is deformed, were true, then it is probable that the condensations mentioned take place only partially or not at all. The cycloketone contg 15 atoms (with the com. name of hexalthone) was mixed with BrH (equimol quantities) and condensation was attempted with cold, dil KOH, hot aq KOH (as with cyclohexanone) and with alc. KOH. There was no condensation in any case. Moreover as in expts of Wieland (no reference) hexalthone (1 mol) and then (cooling with ice) LiNO_3 (1 mol) were added to EtOK (1 mol) in abs EtOH, and the next day H_2O was added but no salt sep., and upon the addn. of water all the hexalthone was recovered. This behavior makes it probable that there is a steric hindrance at work, which would in turn support the hypothesis suggested. Very probably the same arguments apply to polymethylene lactones, e. g., ambrettolides, which have been studied by Kerschbaum. The exptl. details of the present paper will be published later. C. C. DAVIS

Aqueous liquors from low temperature carbonization of coal (MORGAN, PETTET) 21. The protective action of some antioxidants II. The metal halide compounds of some protective agents against aging (KARACIMOV) 30. AcOH recovery in the cellulose acetate industry (CLOTWORTHY) 23. The ultra violet absorption spectra of the amides

of α methylbutenoic acids (CASTILLE) 3. Color and constitution from the standpoint of recent electronic theory (HODGSON) 3. Studies in Raman effect X The Raman spectra of organic substances (DADIEU, KOHLRAUSCH) 3. Raman effect and constitutional problems 11 Cyanogen compounds (DADIEU) 3. Thermal disaggregation of gelatin (GERNGROSS *et al*) 2. Kinetics of heterogeneous formate formation (BIRSTFELT, LOBANOW) 2. Influence of the strength of aromatic amines on their reactions (HERTEL, SCHNEIDER) 2. An x ray study of mannitol, dulcitol and mannose (MARWICK) 2. The crystallographic character of the amides of methylbutenoic acids (THOREAU) 2. Zero volumes of crystalline organic substances (BILTZ) 2. Isoprene and rubber (STAUDINGER, JOSEPH) 30.

ALTFETER, JULIUS Das Hexamethylenetetramin und seine Verwendung. Halle (Saale) W Knapp 178 pp M 14 50 Imco, M 16

BARGELLINI, G Lezioni di chimica organica per gli studenti di medicina e farmacia. Rome Universita 354 pp

BARGELLINI, G Lezioni di chimica organica per gli studenti d'ingegneria. Rome: Universita 123 pp

MEURS, G J VAN Beginselen der Scheikunde. 2. Organische Scheikunde. 3rd ed, revised by H Ph Baudet. Rotterdam Nijgh & van Ditmar. 155 pp. Fl 1 80, bound, Fl 2 25

SMYTHSON, J L The Terpenes. Cambridge Univ Press 420 pp 25s.

Organic compounds. 1 G FARBERND A G (Otto Nicodemus and Walter Berndt, inventors) Ger 514,174, Sept 16, 1925 Addn to 481,819 (C A 24, 2139). Org compounds are obtained by leading the vapors of mixed aromatic heterocyclic ketones, optionally in the presence of oxidizing agents and catalysts, over highly porous materials. Thus, 2-xylyl-5-xylopyridine is evapd under reduced pressure in an air current and passed over active C at 400–420°. The substance 1-xylyl-7-methylbenzoquinoline (m 170°) is found among the unchanged starting material and other products. Other examples are given.

Hydrocarbons. 1 G FARBERND A-G Fr 695,125, May 6, 1930 Products of higher b p are prepd by treating satd. compds. of the aliphatic or hydroaromatic series, particularly the halogen derivs, with hydrocarbons of the C_2H_4 or C_2H_2 series under pressure and in the presence of condensing agents. Thus, hydrocarbons distg between 120° and 200° are prepd from C_2H_4 and $MeCl$ in the presence of $AlCl_3$ and a product distg between 150° and 180° contg about 2% of O from the double compd. of Ft_2O and BF_3 under pressure of C_2H_4 . Cf. C. A 25, 1536 and following abstrs.

Hydrocarbons. 1 G FARBERND A-G Fr 37,296, Aug. 1, 1929. Addn. to 663,539 In the process of the prior patent the conditions of working in the second or following phases are made more vigorous otherwise than by a rise of temp., e. g. by an increase of pressure, an increase of the partial pressure of H, the use of a more active catalyst, etc., or these measures are used along with a rise of temp. (cf. C. A. 24, 2575)

Hydrocarbons. 1 G FARBERND A-G Fr. 37,440, Sept 17, 1929. Addn. to 659,583 (C A 23, 5312) The S and O compds are eliminated from the starting materials by heating these in the presence of catalysts and, if necessary, gases other than H or treating the starting materials with appropriate solvents or in any other manner.

Diolefins. PAUL FEILER (to 1 G Farbernd A-G). U. S. 1,795,549, March 10 In the recovery of diolefins such as butadiene from various mixed gases contg the diolefins, the gas is treated with a salt of a heavy metal from groups 1 and 2 of the periodic system such as Cu_2Cl_2 or $AgNO_3$ and the diolefins are subsequently liberated (suitably by heating) from the resulting additive compds

Organic oxygen compounds. 1 G FARBERND A-G Fr. 37,351, Aug 16, 1929. Addn. to 646,087 (C A 23, 2186) Benzene hydrocarbons having a lateral chain bound cyclically to the C_6H_5 ring, particularly benzopolymethylenic compds. are oxidized by O or gases contg O in the presence of catalysts, preferably heavy metals. Thus, tetrahydronaphthalene is oxidized to α -ketotetrahydronaphthalene and α -hydroxytetrahydronaphthalene. Several examples are given.

Oxygenated organic compounds from carbon oxides and hydrogen. H. DREYFUS. Brit. 338,854, July 24 1929 Products such as EtOH, PrOH, acetaldehyde, HOAc, MeOAc and their higher homologs, and ketones, are formed from mixts contg. C oxides and H such as water gas, coke-oven gas or producer gas in the presence of catalysts consisting of alkali or alk. earth or Cu silicates, borates or phosphates or salts of other oxyacids of P. Temps. of 200–600° and pressures of 50–500 atm may be used,

CH_4 or inert gases such as N may be present and reaction chambers of Cu or steel lined with Cu or of steel contg Mo, W, Mn, Co or Ni may be used. Cf. C. A. 25, 1839.

Mixtures of hydrogen and carbon monoxide. MARIA CASALE SACCHI Austrian 121,228, Sept 15, 1930. Hydrocarbons are treated with O and an excess of steam at a temp above 1000° under reduced pressure.

Oxidizing organic substances. CONSORTIUM FOR ELEKTROCHEM IND G M B H (Erich Baum and Willy O Herrmann inventors) Ger 518,391, May 7, 1921. A mixt of the org substance to be oxidized and an oxygen-contg gas or other oxidizing agent is caused to flow rapidly through a narrow orifice and impinge on a heated surface, which may consist of, or contain a catalyst. The oxidation of AcH , C_2H_5 , and MeOH to $\text{C}_2\text{H}_4\text{O}$, of C_2H_5 to maleic acid and of C_6H_5 to phthalic acid or phthalic anhydride is referred to. Examples are given.

Recovering organic substances from adsorbent materials. I G FARBENIND A G Fr 695,123, May 6, 1930. Org substances are recovered from adsorbent materials contg them by heating them with water, or aq solns or other org or inorg liquids in which the org substances are scarcely or not at all sol in the presence of substances which are more easily wetted by the org substances than the adsorbent materials which contain them. Thus fuller's earth contg oil is mixed with active C and soft soap and the mass is boiled with water. An oily layer contg the C seps, and the C is removed by filtration.

Hydrocyclic compounds. I G FARBENIND A G Fr 37,498, Oct. 9, 1929. Addn to 663,564 (C A 24, 625). Hydrocyclic compds are prepd by condensing unsatd $\alpha\beta$ nitriles contg the atomic group $-\text{C}=\text{C}-\text{C}=\text{N}$ in the form of a chain, with

compds contg doubly conjugated $\text{C}=\text{C}$ compds. Thus, the nitrile of crotonic acid ($\text{CH}_3\text{CH}=\text{CH}-\text{CN}$) is heated under pressure with cyclopentadiene. The product has

probably the formula $\text{CH}_2=\text{CH}-\text{CH}=\text{CH}-\text{CH}(\text{CN})-\text{CH}_2$. With 2,4-dimethyl-

butadiene, a product of the formula $\text{CH}_3-\text{CH}(\text{Me})-\text{CH}(\text{CN})-\text{CH}(\text{Me})-\text{CH}_2$ is obtained. Products from the nitrile of sorbic acid are also described.

Basic products. I G FARBENIND A G Fr 37,289, July 30, 1929. Addn to 671,362 (C A 24, 2244). Imido ethers, the acids from which they are derived, their chlorides, esters or amides are condensed with primary or secondary aromatic amines contg a basic lateral side chain with a tertiary N, or the side chain with a tertiary N may be introduced afterward. Thus, if β -diethylaminoethoxyaniline is condensed with oleic acid chloride a product of the formula $\text{C}_{17}\text{H}_{33}\text{CONHC}_2\text{H}_5\text{OC}_2\text{H}_5\text{NEt}_3$ is obtained. Other examples and the formulas of the products obtained are given. Cf. C A 25, 837, 1261.

Diaryl compounds. FEDERAL PHOSPHORUS Co Fr 37,449, Sept 20, 1929. Addn to 667,840 (C A 24, 1394). Complex diaryl compds of b p higher than biphenyl are produced by vaporizing the corresponding hydrocarbon, e. g., pure com benzene or com 90% benzene contg aromatic hydrocarbons other than $\text{C}_{12}\text{H}_{10}$, such as toluene and xylene. The preliminary heating is to a temp just below that at which diaryl compds are formed and the temp is then rapidly raised to that at which these compds are formed.

Complex diaryl compounds. J N CAROTHERS, T. J SCOTT and FEDERAL PHOSPHORUS Co Brit 338,631, Sept. 5, 1929. See preceding abstr.

Drying organic liquids. WILHELM KOPESKA Austrian 120,875 Aug 15, 1930. Benzene and other org liquids contg water in suspension are dried by treatment with a higher fatty acid ester of a carbohydrate of the formula $(\text{C}_{12}\text{H}_{21}\text{O}_2)_n$. The ester must be insol in the org liquid, and may be used in such a form that the org liquid can filter through it. Cellulose stearate, palmitate and laurate are suitable esters.

Trichloroalcohols. CHEMISCHE FABRIKEN JOACHIM WIERNIK & Co A G (Gustav Heidner, inventor) Ger 515,539, Jan 14, 1925. The corresponding ketone is caused to react with CHCl_3 in the presence of alkali amide, and the resulting alcoholate treated with acid. The alkali amide may be suspended in a water free indifferent solvent. Thus, acetone and CHCl_3 are added to a suspension of NaNH_2 in Et_2O . The temp is kept at 0° . Dil H_2SO_4 (33%) is added and a steam distn gives a 60% yield of $\text{C}_3\text{H}_5\text{Cl}_3$. Further examples are given.

Hydrogenating paraldol. I G FARBENIND A G Brit. 338,807, Jan 30, 1929. 1,3-Butylene glycol is made by catalytic hydrogenation of paraldol at $30-50^\circ$ in the presence of a solvent or diluent such as water, EtOH , isopropanol or butanol, under

pressure, with catalysts such as Ni or Cu or both on pumice or kieselguhr Use of Co, Pt and Pd also is mentioned

Ketones such as acetone. H DAEVUS Brit 338,518, Aug 19, 1929. The vapors of aliphatic alcs contg at least 2 C atoms, in admixt with air or O, are subjected to temps of 250–700° in the presence of catalysts such as oxides, hydroxides, carbonates, metaborates, ortho-, meta- and pyro-phosphates and silicates of Ca, Ba or Mg which may be deposited on pumice or kieselguhr Natural compds such as wollastonite, augite, osteolite, danburite or datolite may be used, and the gases may be passed through a tube of Cu, Fe, staybrite or earthenware contg the catalyst and the vapor of an aliphatic acid such as HOAc may be added and a large excess of O used Brit 338,519 relates to a similar process in which steam is added to the reaction mixt

Carbocyclic ketones Brit 339,348, Oct 9, 1928, Soc ANON M NARF & CIE. Fr 37,380, Aug 30, 1929 Addn to 599,765 Unsatd cyclic ketones having more than 9 ring members are prepd by submitting to the action of heat a metallic salt of an unsatd dicarboxylic acid contg at least 11 atoms of C in the normal chain and the 2COOH groups at the ends of the chain Salts of UO_2 or metals having the at numbers 59–72 are referred to

Methylene ethers. IMPERIAL CHEMICAL INDUSTRIES, LTD, T BIRCHALL and S COFFEY Brit 338,624, Aug 30, 1929 In the production of products such as methylene diethyl ether methylene di *n* butyl ether and 1,3 dioxane, $(CH_3)_2N_2$ is caused to react with mono- or di hydric alcs in the presence of acids such as HCl or H_2SO_4 Among the alcs which may be used for such reactions are MeOH, EtOH, *n* Bu alc, allyl alc, benzyl and substituted benzyl alcs, and ethylene and trimethylene glycols

Diazo compounds from 4-aminoresorcinol. N N VOROZHTOV Russ appl 49 050, June 12, 1929 4,1,3 $H_2NC_6H_3(OH)_2$ is diazotized by known methods in the presence of Cu or Fe salts or Co, Ni, Zn or other similar metal

Diazoamino compounds. I G FARBENING A-G Fr 37,240, July 25, 1929 Addn to 677,579 (C A 24, 3248) Diazoamino compds are made by the reaction of diazo compds on polycyclic primary amines sol in water, which on account of the position of the substituents, with respect to the amino group, are not capable of forming azo dyes Thus, 3 nitro-4-amino-1-methylbenzene is diazotized and combined with 1-naphthylamine-2,4-disulfonic acid Other examples are given

Aromatic amines. I G FARBENING A-G (Johus Laux, inventor) Ger 515,758, Aug 21, 1925 Aromatic amines are prepd by reducing the corresponding nitro compds with Fe in the presence of aq Al salt solns Thus $PhNO_2$ is reduced to $PhNH_2$ by Fe in the presence of aq $AlCl_3$ After distg off the $PhNH_2$, a yellowish brown ppt is left, which, on sepu from the unused Fe, produces a deep violet red Fe dye on heating to glowing Further examples are given

N-Aminoalkylamines. I G FARBENING A-G (Werner Schulemann, Fritz Schönhofer and August Wingler, inventors) Ger 518,207, Jan 26, 1927. Amines are caused to react with aminoalkyl halides or the corresponding sulfonates, the amino-alkylating agents being used in the form of their salts, in the presence or absence of an acid binding agent and a solvent Thus, ω -diethylaminoethylamine may be prepd by heating $PhNH_2$ with $Et_2NCH_2CH_2Cl$ HCl Other examples are given also Cf C A 24, 2469

Esters. E I DU PONT DE NEMOURS & Co Ger 515,678, Mar 25, 1928 Esters and other org compds result when primary alcs with more than 1 C atom or a mixt of alcs are subjected to the action of dehydrogenating catalysts at raised temp and pressure. The alcs are circulated over an equal amt of catalyzer, volume-for-volume, at a pressure of 10 atm The catalysts may be Ni, Cu, Co, Fe alone or mixed with MnO_2 , Cr_2O_3 , MgO, CaO, etc. EtOH treated in this way gives 55% unchanged EtOH, 25% AcOEt, 12.5% BuOH and small quantities of AcOH, AcEt, CO_2 and CH_4

Lactic acid esters. SCHERING-KAHLBAUM A-G (Frank Hartwich, inventor) Ger. 518 388, Jan 23, 1926 A mixt of Mg lactate, an alc. and an acid, with or without a catalyst, e g., an excess of acid or $CaCl_2$, is heated Examples are given

Halogenated phosphorous and phosphoric esters. I G FARBENING A-G Brit 338,981, Aug 31, 1929 Reaction is effected between a compd contg an alkylene oxide group and a P trihalide or oxyhalide, suitably with use of a catalyst such as I, $FeCl_3$, S or iron filings E g., with PCl_3 and ethylene oxide, the O bridge is opened and a Cl atom combines with a CH_2 group while the P atom combines with the O. Thus ethylene oxide can be attached up to three times, so that the neutral ω -chloroethyl ester of phosphorous acid may be formed Examples are also given of reactions between epichlorohydrin and P oxychloride with iron filings as catalyst and of epichlorohydrin with PBr₃

Alkyl halides. N-V DE BATAAFSCHE PETROLIUM MAATSCHAPPIJ Brit. 338,742, Jan 8, 1929 Products such as pentyl, butyl and propyl chlorides and butyl bromide are made by chlorination or bromination of paraffin hydrocarbons such as propane, butane and pentane, as by mixing the reacting materials in the dark at a low temp. and then heating without exposure to light (suitably in app. of enameled metal or Pyrex glass)

Addition products from acetylene and hydrogen halides. CONSOATIUM FOR ELEKTROCHEMISCHE INDUSTRIE GEZ. Brit. 339,093, Nov 19, 1928 Addn. products are formed by passing a mixt. of C_2H_2 and a halide such as HCl or HBr at a raised temp. over a catalyst comprising a compd. of an alk. earth metal Mg, Bi, Sb, V, Zn, Al, Fe or lig. or a mixt. of such compds., carried on a material of large surface such as active C or SiO_2 gel

1-Aroylamino-4-haloanthraquinones. I G FARBENIND A-G (Willy Hartmann, inventor) Ger. 518,406 Dec 23, 1926 1-Aroylaminoanthraquinones are halogenated in inert solvents. Examples are given

Anthraquinoneazines. W SMITH S G WILDMOTT, J THOMAS and SCOTTISH DYES, LTD Brit. 338,486 May 13, 1929 Anthraquinonetetrahydrodiphenazines are made by condensing α,α -dichloroanthraquinones with *o*-nitroaniline or α,α -diaminoanthraquinones with *o*-nitrochlorobenzene, and treating the products with alc. Na sulfide. The condensation may be effected by heating in an inert solvent such as $PhNO_2$ in the presence of an acid-absorbing agent such as Na_2CO_3 and a Cu catalyst, and the hydrodiphenazines may be converted into phenazines by treatment with oxidizing agents. Several examples are given

Mercaptobenzothiazoles. I G FARBENIND A-G (Eduard Tschunkur and Ernst Herdickehoff, inventors) Ger. 518,206, Feb 22, 1928 See Fr. 669,921 (C. A. 24, 1867)

Monoxanthogens. GEORGE S WHITBY (to Roessler & Hasslacher Chemical Co.) U.S. 1,796,972, March 17 Monoxanthogens of the general formula $RO-CS-S-CS-OR$, wherein the group $RO-$ is the radical of any aliphatic alc. or deriv. thereof capable of yielding a xanthate and wherein R is of a non-cyclic nature can be prepd. in excellent yield by the reaction of alkali metal cyanides with the corresponding disulfide compds. as follows: $RO-CS-S-CS-OR + MCN = RO-CS-S-CS-OR + MSCN$ (where M is an alkali metal). The disulfides used may be prepd. by a variety of methods, such as electrolysis of concd. xanthate solns., action of I with KI on a xanthate soln., etc., but preferably are prepd. through the action of an alkali metal hypochlorite on a xanthate. The propyl, butyl, isobutyl, isoamyl and hexyl derivs. are yellow oils, while the methyl deriv. forms yellow needles m. 55° . The monoxanthogens prepd. from a mixt. of higher alcs. (corresponding in mean mol. wt. to decyl alc.) as well as from the monoethyl ether of ethylene glycol, by way of the corresponding xanthates and dioxanthogens, are also yellow oils. Those monoxanthogens in which $R-O-$, of the formula $R-O-CS-S-CS-O-R$ is a secondary aliphatic alc. radical, are good accelerators of the vulcanization of rubber and are called secondary alkyl monoxanthogens. As an example of this group the isopropyl deriv. or isopropyl monoxanthogen, is a golden yellow cryst. solid m. $54-55^\circ$. Details are given of the production of isopropyl monoxanthogen.

Calcium humate. R KÉKÉRS. Hung. 100,679, Oct. 24, 1928 Materials in the state of primary carbonization (e.g., peat, lignite or brown coal) are finely powdered and boiled under pressure with alkalis. The alkali humates are filtered and lime water is added to ppt. Ca humate.

Anthraquinone derivatives. I G FARBENIND A-G. Fr. 37,109, June 17, 1929 Addn. to 630,511 Esters of *o*-aminocarboxylic acids of the anthraquinone series are prepd. by treating an anthraquinone 1,2 isoxazole with a compd. of the formula ROH in which R is an alkyl or aryl group such as alcs. or phenols in the presence of an agent of alk. action. Examples are given of the prepn. of the Et, Me and Ph esters of 1-aminoanthraquinone-2-carboxylic acid.

Azine derivatives. I G FARBENIND A-G. Brit. 339,283, Aug 30, 1929 1,2,5,6-Dibenzophenazine and its derivs. are made from 2-naphthylamine or its corresponding derivs. which contain a free 1 position by oxidation in an inert solvent of high b.p. with a metal oxide such as CuO or MnO_2 . Several examples are given. Cf. C. A. 24, 2242.

Benzanthrone derivatives. I G FARBENIND A-G (Otto Braunsdorf, Eduard Holzapfel and Hans Lange, inventors) Ger. 518,410, Aug 3, 1926 Addn. to 479,356 (C. A. 23, 4830) The ammobenanthronyl sulfides, etc., described in Ger. 479,356 are now prepd. by the action of NH_4OH in H_2SO_4 soln. on benzanthrone sulfides, bisulfides, mercaptans or their derivs. The reaction is effected at about 125° , and a

catalyst may be present, e. g., FeSO_4 or CuSO_4 . Examples are given. Cf. C. A. 25, 2301.

Cyclohexanol derivatives. SCHERING-KAHLBAUM (Walter Schoeller and Hans Jordin, inventors). Ger 514,417, July 24, 1927. Addn to 512,719 (C. A. 25, 1260). Alkyl- α -alkylcyclohexanol and corresponding ketones are prepd by subjecting a mixt of the corresponding phenol and alkylated coumarin to the action of H, in the presence of catalysts, until 6 or 8 H atoms are taken up, and sepg the alkylated coumarin from the hydrogenated product. Thus a mixt of 4-methyl-6-isopropylphenol and 2,4-dimethylcoumarin is treated with H at 150° in the presence of a Ni catalyst. Pressure may be applied. The reaction is stopped when 8 atoms are taken up. A 60% yield of the hydrogenated product *p*-menthol and a 3-5% yield of *p*-menthone, can be obtained by careful fractional distn. Further examples are given.

Guanidine derivatives. REINHARD FREIHERR VON GODIN. Ger 518,407, July 24, 1928. Salts of diguanidine derivs of higher paraffin hydrocarbons are prepd by the reaction of α,ω -dihalogenated higher paraffin hydrocarbons with guanidine in warm alc. soln. About 3 mols of guanidine should be used for 1 mol of halogenated hydrocarbon. Examples are given.

Substituted guanidines. HERBERT SCHOTTE (to Schering-Kahlbaum A. G.). U. S. 1,795,738, March 10. In forming a compd such as α -dimethyl- β -ethylguanidine, a dialkylcyanamide such as dimethylecyanamide is treated with a salt of an amine having at least one H atom attached to the N atom, such as ethylamine-HCl in the presence of the corresponding free amine.

Diacyl derivatives of *m*-xylene. I. G. FARBERIND A. G. (Otto Wulff, inventor). Ger 515,540, Dec 21, 1926. Aromatic or aliphatic acid chlorides are caused to react on monoacyl *m*-xylene derivs. or their AlCl_3 compds at temps above 50°. Excess of AlCl_3 and diluents may be present. In an example, acetyl-*m*-xylene is mixed slowly with AlCl_3 at 40°. AcCl is dropped in at about 45°. The temp is then raised to 85-90° for two hours, i. e., until the HCl evolution ceases. The temp is then raised to 95°. The reaction mass gives an 80% yield of diacetyl-*m*-xylene. Further examples are given.

Reduction products of *N*-aretylindoxyl. I. G. FARBERIND A. G. (Curt Schumann, Eduard Münch and Bruno Christ, inventors). Ger 515,544, May 27, 1928. *N*-Aretylindoxyl, its homologs, analogs or derivs. are treated with H in the presence of hydrogenating catalysts. Numerous examples are given. In one, *N*-aretyl-2,3-naphthindoxyl is treated with H at 100 atm at 25° for 16 hrs., in an autoclave, in the presence of alc. and a Ni kieselguhr catalyst, to give *N*-aretyl-2,3-naphthodihydroindoxyl. Cf. C. A. 24, 2142.

Solid compounds from hydrogen peroxide combinations. I. G. FARBERIND A. G. Brit 339,332, Aug 6, 1929. Gaseous mixts. contg vapors of H_2O_2 are passed (suitably countercurrentwise) over substances which unite with it such as urea, methylolurea, acetamide, urethan, succinic amide, mannitol, erythritol, pinacol, $(\text{CH}_3)_2\text{N}_2$, or alkali metal sulfates, carbonates, borates, nitrates, phosphates or pyrophosphates (either with cooling, heating (to not above 80°) or at ordinary temp). Stabilizers such as starch, dextrin, amines, EtOH , borates or acid or neutral pyrophosphates may be incorporated in small proportions with the solid substances, and the gaseous mixt used may be obtained by the action of a silent elec. discharge or of ultra violet rays on H contg air or O_2 .

Concentrating acids. SOC. ANON. DES DISTILLERIES DES DEUX-SÈVRES. Fr 77,269, Mar. 19, 1929. Addn to 651,528 (C. A. 23, 3234). The processes of the prior patent and its addns. 34,574 (C. A. 24, 1392) and 34,601 (C. A. 24, 1392) are applied to the dehydration of org. acids in general and of the fatty series in particular. A solvent insol or slightly sol in water is used the b. p. of which is higher than that of the acid to be dehydrated, and which does not form an azeotropic mixt. with the anhyd. acid. A carrying substance giving with water an azeotropic mixt. boiling below the mixt. of water and solvent is also used. The method described in the addn. 35,750 (C. A. 24, 4052) is varied by using an accessory liquid which forms with the acid to be dehydrated a binary mixt. of min. b. p. Cf. C. A. 25, 970.

Acylated hydroxy fatty acids. ORANIENBURGER CHEMISCHE FABRIK A. G. Ger. 515,679, Aug 24, 1928. Hydroxy fatty acids or derivs. are treated with org. acids or their anhydrides or chlorides. A small quantity of CISOH is added during the acylation, and the product is neutralized after the reaction is over. Thus, ricinoleic acid is mixed with BrCl and a small quantity of CISOH added. After the reaction is over, the product is neutralized with NaOH . Further examples are given.

Mixtures of hydroxy fatty acids. CARL STIEFEL. Ger 518,300, June 3, 1929.

Acid mixts resembling the acids of castor oil are prepd from unsatd. animal or vegetable fats or oils, or their free acids or their soaps, by partially satg these by chlorination, and then replacing the Cl atoms by OH groups in known manner, by heating with NaOH soln under pressure, etc. 1 sample is given.

Acid- and lime-resisting derivatives of unsaturated fatty acids. FREDERICK MONZ (to General Aniline Works) U. S. 1,759,801, March 17. Acids such as crotonic acid or nicotinic acid (heated to 150-170° until its acidity is reduced by about one-half, with formation of dimeric acid) are treated with fuming sulfuric acid contg 65% SO_3 (suitably in CCl_4) to obtain products which are not pptd from soln by hard water and are resistant to acids.

Pyrazoloanthrone-2-carboxylic acid. I. G. FARBEN- und A. G. (Heinz Scheerer, inventor) Ger. 515,640, Apr. 6, 1929. NaCl_2 is caused to react on 1 haloanthraquinone-2-carboxylic acid. 1 sample is given. Cf. C. A. 25, 909.

Phenylglycidic acid. ADOLF LAUFMANN Ger. 515,031, Oct. 30, 1928. Addn. to 509,938 (C. A. 25, 714). The method of 509,938 for forming the above acid by oxidation of cinnamic aldehyde is modified by using hypohalogenite soln as the oxidizing agent. Examples are given.

1-Hydroxy-4-haloanthraquinone-2-sulfonic acids. I. G. FARBEN- und A. G. (Ernst Herold, inventor) Ger. 519,214, Aug. 29, 1929. The diazo compds from 1-amino-4-haloanthraquinone-2-sulfonic acids in an aq. medium are treated with Cu or a Cu compd in the presence or absence of free acid. The amt of Cu may be such that the product seps as the insol. Cu salt. Examples are given.

Formic acid. N. V. DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ Brit. 339,045, Oct. 31, 1928. See Fr. 653,688 (C. A. 24, 4522).

Acetic acid. HERMANN SUDA, Austrian 130,879, Aug. 15, 1909. Conc'd AcOH is recovered from dil. AcOH by extrn. in the liquid phase with esters b. above 150° and derived from mono- or poly basic cyclic carboxylic acids and mono- or poly hydric alcs. Examples describing the use of dibutyl and diethyl phthalates are given. Cf. C. A. 24, 3250.

Concentrating acetic acid. HERMANN SUDA U. S. 1,705,977, March 10. Dil. HOAc is subjected to cold extrn. with a solvent such as anthracene oil and quinoline and the acid and extg. medium are sep'd by distn.

Concentrating acetic acid. HERMANN SUDA U. S. 1,701,009, March 10. A superheated aq. HOAc vapor mixt. is brought into direct contact with a counter flow of a liquid ester such as dibutyl phthalate to effect extrn. of the HOAc from the mixt. and conc'd HOAc is recovered from the soln. thus formed. App. is described.

Recovering acetic acid. HERMANN SUDA, Austrian 121,251, Sept. 15, 1920. Dil. AcOH is extrd. in the cold with a mixt. of a non-acetylizable base b. above 150° and pentachloroethane or a hydrocarbon b. above 150°. A suitable extg. mixt. is quinoline 1 and anthracene oil 1 part. Conc'd AcOH is recovered by distg. the ext.

Acetic acid and other acids from nitriles. I. G. FARBEN- und A. G. Brit. 339,235, Dec. 18, 1929. Nitrile vapors, either directly as produced by the method described in Brit. 322,258 (C. A. 25, 115) or after previous condensation and revaporization, are saponified by continuously passing the nitrile vapor in contact with a current of more acid such as a 50-99% H_2SO_4 or H_3PO_4 . Various details and examples are given.

Anhydrides of aliphatic acids. N. V. DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ Brit. 338,556, Aug. 9, 1929. Anhydrides such as Ac₂O or butyric acetic anhydride are obtained by heating monobasic aliphatic acids in admixt. with monoketones such as acetone or ketone forming substances such as iso-Pr alc., secondary Bu alc., propylene oxide or MeOAc (suitably at about 600-700° and under ordinary, increased or reduced pressures). The materials may be passed through a heated tube of porcelain, quartz, yellow brass or other suitable material and catalysts may be used such as Cu, yellow brass or Al sulfate together with fillers such as porcelain, kieselguhr or pumice. Solid ketone forming substances such as NaOAc or Ca(OAc)₂ may be used and caused to pass through the reaction tube by a device such as a screw conveyor.

Dehydrogenating isopropyl alcohol. N. V. DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ Brit. 339,491, Dec. 6, 1929. Isopropyl alc. vapor, with or without admixt. with other gases, is passed through a liquid medium such as molten Pb or salts or other non reactive materials, heated to 300-700°, to produce acetone and H₂. Various details of app. and procedure are described.

Polymers of vinyl alcohol. I. G. FARBEN- und A. G. (Fritz Klatte and Hermann Müller, inventors) Ger. 515,780, Nov. 29, 1929. Addn. to 514,593 (C. A. 25, 2153). Polyvinyl mixed esters, which contain a halogenated org. acid and a non halogenated

ox. acid, are sapon. by NH_4OH . Thus, polyvinyl acetate chloroacetate is sapon. with NH_4OH to give a solid polyvinyl alc.

Acetaldehyde. **CONSORTIUM FÜR ELEKTROCHEM. IND. G. M. B. H.** (Erich Baum and Martin Mugdan, inventors) Ger 518,200, July 22, 1931. An excess of C_2H_2 is led through aq. H_2SO_4 of 6–35% concn contg a compd. of Hg, the acid being maintained at $(40-50)^\circ$. ACH is removed from the gases leaving the reaction vessel, and the excess of C_2H_2 is returned to the vessel.

Acetylene. **RUSSELL W. MILLAR** to The Shell Development Co.; Can. 309,544, Mar. 17, 1931. C_2H_2 is formed by the controlled combustion of a stream of O-contg gas in an atm. of hydrocarbon gas.

Normal butyl oleate. **W. M. J. RAYNISTER** (to Commercial Solvents Corp.) U. S. 1,796,231, March 10. Normal Bu oleate is made by reaction of butanol and oleic acid in the presence of H_3PO_4 . It is a mobile yellow oily liquid, insol. in water, $b.p.$ is $235-43^\circ$.

Acetals. **WILLY O. HERRMANN** and **HANS DEUTSCH** (to Consortium für Elektrochemische Industrie); U. S. 1,796,190, March 17. See Ger 502,431 (C. A. 24, 479).

2,5,6-Trichloro-1,3-dimethylbenzene, etc. **GEORG KALISCHER** and **FRIEZ FRÄSTER** (to General Aniline Works) U. S. 1,796,108, March 10. 2,5,6-Trichloro-1,3-dimethylbenzene, crystg. from glacial HOAc as white needles m. about $95-6^\circ$, b. $255-60^\circ$, is formed by the action of an inorg. acid on the corresponding 4-sulfonic acid (which in turn is obtained by reaction of NaOCl on 1,3-dimethylbenzene-4-sulfonic acid in aq. HCl soln. and in general chlorinated products are formed by treating 1,3-dimethylbenzene-4-sulfonic acid in a mineral acid soln. with a chlorinating agent. Depending on the amt. of Cl employed the mono-, di- or tri-chloro-1,3-dimethylbenzene-4-sulfonic acid is obtained. By hydrolysis effected in the usual manner (for instance, by heating with H_2SO_4 or H_3PO_4 or the like, the sulfonic acid group is split off from the mol. of the aforesaid compds. and the corresponding chlorinated benzene derivs. are obtained. For splitting off the sulfonic acid group from the trichloro-1,3-dimethylbenzenesulfonic acid an excess of Cl may be employed, whereby the sulfonic acid group is replaced by C and tetrachloro-1,3-dimethylbenzene is obtained. Details of procedure are also given for the production of 2-chloro-1,3-dimethylbenzene-4-sulfonic acid (the amide of which m. about 223°), 2-chloro-1,3-dimethylbenzene (b. about $185-7^\circ$), dichloro-1,3-dimethylbenzene-4-sulfonic acid (the amide of which m. about 330°), tetrachloro-1,3-dimethylbenzene (m. $219-20^\circ$) and derivs.

m-Nitroaniline. **G. A. KIRCHOV** and **I. T. EAKIN** Russ. appl. 66,606, Mar. 18, 1930. $m\text{-ONC}_6\text{H}_4\text{NH}_2$ is prepd. by reducing an aq. suspension of $m\text{-C}_6\text{H}_4(\text{NO}_2)_2$ with Na_2S in the presence of NH_4Cl .

2,4-Dihydropyridine-3-carboxylic nitrile. **GEORG SCHROETER**, Ger. 515,991, Jan. 18, 1927. The above nitrile and its derivs. are prepd. from dihydroxyhalopyridine carboxylic nitrile by (a) replacing the halogen by H or a substituted NH_2 group, or (b) treating with 1 mol. of an alkylating agent to form the N-alkylated pyridone compd., and substituting the halogen by H, OH, SH, NH_2 or substituted NH_2 . Thus, 2,4-dihydroxy-6-chloropyridine-3-carboxylic nitrile is treated with 2 N H_2SO_4 and Zn dust to give the required nitrile. Or, the chloro deriv. is treated with NaOH and Me_2SO , to give 1-methyl-4-hydroxy-6-chloro-2-pyridone-3-carboxylic nitrile. The Cl is then substituted by H and treated with H_2SO_4 and Zn as before. Further examples are given.

Pure phenylazo-2,6-diaminopyridine hydrochloride. **C. F. BOEHRINGER & SOHN** G. M. B. H. Ger 515,781, Nov. 5, 1927. 2,6-Diaminopyridine is coupled with diazotized PhNH $_2$ in the presence of a sufficient amt. of dil. HCl to form the dihydrochloride at 12° . The dihydrochloride is then decomposed to the monohydrochloride by addn. of water. An example is given.

Hydroxythionaphthenes. **KARL SCHIRMACHER** and **ERNST FISCHER** (to General Aniline Works) U. S. 1,797,104, March 17. Hydroxythionaphthenes can be obtained

with an excellent yield from compds. of the formula: $\text{CX} \cdot \text{CX} \cdot \text{CX} \cdot \text{C alkyl} \cdot \text{CY} \cdot \text{CS} \cdot \text{CH}_2\text{CO}_2\text{Z}$ wherein X may be H, alkyl or halogen, Y may stand for CN, COOH, COO-metal and Z for H or a metal, by heating them with water to a temp. of $100-200^\circ$, preferably 180° , that is to say, by using such α -carboxythioglycolic acids, or salts thereof, or α -cyanothioglycolic acids as contain in α -position to the carboxyl or nitrile group a further substituent, particularly an alkyl residue. If there is no further substituent in α -position to the cyanogen group, no hydroxythionaphthene is formed. Examples with details of procedure are given for the production of 4-methyl-6-chlorohydroxythionaphthene and 4,7-dimethyl-6-chlorohydroxythionaphthene. Cf. C. A. 25, 973.

Alkoxy-3-hydroxythionaphthenes. I. G. FARBENIND A-G (Frist Runne, Karl Moldaenke and Ernst Fischer, inventors) Ger 515,541, Oct. 9, 1927. Alkoxyaryl thioxycolic acid halides of the $C_{10}H_8$ or $C_{11}H_9$ series are treated with Al halides in the presence of solvents. Thus, 6-ethoxynaphthalene 2 thioxycolic acid (m 122°) in C_6H_5Cl is converted to the chloride by addn of PCl_5 . $AlCl_3$ is added at 20-40° and the C_6H_5Cl steam-distd off. An 80% yield of the corresponding ethoxybenzohydroxythionaphthene is obtained. Further examples are given. Cf C A 24, 370.

4-Alkylquinolines. SCHIFFING KAHNBAUM A-G (Clemens Zöllner, inventor) Ger 518,291, Jan 15 1927. $PhNH_2$ or its substitution products are caused to react with alkyl β haloethyl ketones in an acid or alk. soln. in the presence of an oxidizing agent. Thus 4-methylquinoline is prepd by heating on the water bath a mixt of $PhNH_2$ with methyl β chloroethyl ketone, concd HCl , and $PhNO_2$. Other examples are given also.

Anthraquinone and its derivatives. I. G. FARBENIND A-G (Er 37,337, June 26, 1929. Addn to 173,825 (C A 24, 2757). The transposition products, obtained by treating with alk. or acid agents or by heating the products resulting from the condensation of one mol. of benzoquinone with 2 mols., identical or different, of 1,3-butadiene hydrocarbons are submitted to a regulated dehydrogenation which may be followed by a more vigorous dehydrogenation. The transposition and the regulated dehydrogenation may be carried out in one operation. 1,4,5,8-Tetrahydroanthrahydroquinone, 1,4,5,8-tetrahydroanthraquinone, their homologs and derivs. prepd by this process are stated to be new. Cf C A 25, 965.

2-Aroylpyrazoleanthrones. I. G. FARBENIND A-G (Heinz Scheyer, inventor) Ger 518,120, Apr 6 1928. Pyrazoleanthrone 2-carboxylic acid or its *N*-alkyl or *N*-aralkyl derivs. or substitution products of any of these, are converted into the chlorides, and these are condensed with aromatic hydrocarbons or their substitution products in the presence of an acid condensing agent. Examples are given.

***N*-Substituted 3,6-diaminoxanthenes.** IMPERIAL CHEMICAL INDUSTRIES. Ger 515,524, Mar 3 1929. See Brit 314,806 (C A 24, 1521).

Phenols. F. FEARLHOFF. Brit. 328,628, Sept 11, 1929. Phenols are sepd from phenolic tar oils by treating the oils with aq. solns. of alkali carbonates under pressure and at temps. above 100°, with continuous escape of CO_2 , isolation of the phenates thus formed and then decomn. of them with CO_2 liberated in the process, to regenerate the phenols. Various details of procedure and app. are described.

α -Naphthol. I. G. FARBENIND A-G (Friedr. Laage, inventor). Ger 518,409, Apr 11, 1929. See Ger 503,506 (C A 25, 1529).

Coumarone. GES. FÜR TREIBERBEREITUNG M. B. H. and OTTO KÄLNER. Ger 515,546 July 14 1929. Heavy benzene contg. coumarone, is sulfonated and cooled in the presence of at least sufficient water absorbing agent (e. g., Ac_2O) to absorb the water formed during sulfonation. The resulting sulfonic acids are then split up by H_2 , causing pure coumarone to sep. Examples are given.

Isatins. DEUTSCHE GOLD- UND SILBER-SCHNEIDANSTALT NORMALS ROESSLER. Ger 515,542, May 13, 1927. Isatins are prepd by converting *N*-substituted cyanofarmylides into the corresponding imide compds. at high temps. in the presence of condensing agents, and saponifying the product. Thus, diphenylcyanofarmamide is shaken with concd H_2SO_4 . The product is cooled and treated with hot dil. $NaOH$, to give *N*-phenylisatin m 140°. Other examples describe the prepn. of *N*-methylisatin, and *N*-ethylisatin.

Diphenylisatin. G. RICHTER, LTD. Hung 101,103, Jan 28, 1920. Isatin is condensed with phenol in dil. mineral acids.

1,2,4-Triazoles. GEORG SCHELING and BRUNO WALACH (to C. H. Boehringer Sohn A. G.) U. S. 1,796,403, March 17. Triazoles are in general obtained, e. g., by treating oxime ester, e. g., the Lenzene-sulfonic acid esters of the oximes, with acylhydrazines under conditions under which the esters undergo Beckmann transformation to the imido ester of the corresponding secondary acid amide. Further, triazoles may also be obtained by subjecting oximes to the Beckmann transformation with the aid of acid halides, such as PCl_5 , $POCl_3$, and the like and by causing acylhydrazines to react with the reaction product. The same imido compds. and therefore also the triazoles may be obtained from mono-substituted acid amides by acylating the same in their enolic form, e. g., by treating the mono-substituted acid amides with acid halides, such as acid chlorides, and then causing an acylhydrazine to react with the reaction product. The 1,2,4-triazole is formed by the elimination of water from the resulting primary hydrazidine. Numerous examples with details of procedure are given and particular claim is made to 3-methyl-4-xylyl-5-isopropyl-1,2,4-triazole as a new product, manuf. of which and of various similar compds. is described.

Thymol, etc. SCHERING KAHLEBAUM A-G (Hans Jordan, inventor). Ger 518,209, Aug 17, 1926 Thymol and its isomers and homologs are prepd by hydrogenating the phenol ketone condensation products obtainable as described in Brit 273,684 (C A 22, 1982), the process being continued until 4 atoms of H have been taken up. A phenolate or a compd giving rise thereto may be present in the reaction mixt, as well as a catalyst promoting hydrogenation. Examples are given. Cf C A 25, 717.

Porphyrin. HANS FISCHER and JOSEF KLARER. Ger 515,992, Oct 9, 1926. Addn to 509,939 (C A 25, 716). The method of 509,939, for producing porphyrin by acting on halogenated dipyrromethenes with concd H_2SO_4 , is extended to include other acids. Examples mention HCOOH , succinic acid, AcOH and H_3PO_4 . Cf C A 25, 716.

Skatole. GES. FÜR TEERVERWERTUNG M B H and OTTO KRÜBER. Ger 515,543. Nov 10, 1928. Skatole is obtained by treating the tar-oil fraction contg indole homologs with Na or NaNH_2 at temps above 165° , and sepg the unchanged oil from the Na indole compds. The skatole is then sepd from the oil by direct cooling. The mixt of Na indole compds may also be treated with CO_2 with exclusion of water, to cause the skatole to sep. Thus, crude indole oil is treated with Na at $190-5^\circ$. A stream of dry CO_2 is led in. An oil seps above the lye. This is cooled to a low temp to cause solid skatole to sep. Another example is given.

11—BIOLOGICAL CHEMISTRY

PAUL E. HOWE

A—GENERAL

FRANK P. UNDERHILL

A model for the segregation of enzyme from substrate (diastase-starch) in the cell. M. MICHLEBAUER. *Fermentforschung* 12, 273-94 (1931).—The sepn existing between enzyme and substrate in the living cell and the more or less marked abolition of this sepn under definite conditions is illustrated with pancreatic diastase and starch as a model. This spatial sepn in the model is effected by superficially active substances which occur in the cell, such as Na oleate, lecithin, salts of bile acids and cholesterol. With Na oleate the sepn of enzyme from substrate is brought about by its property of surrounding enzyme and substrate particles and thereby imparting to them a definite and uniform elec charge. The enveloping film is thus a superficially active substance and the enclosed particles are mutually repelled by similar charges. The sepn is reversible under biol conditions. The reversal is best effected by changing the H ion concn. In a mixt of diastase and starch without soap, a change in pH between 6.8 and 8.0 hardly exerts a measurable influence on the rate of digestion. If, however, the system contains soap, a slight change in pH within these limits is of enormous influence. The time required for disappearance of the starch, as shown by neg 1 reaction, may thus be several hrs. at the higher and only a few sec. at the lower pH . The change in pH toward the acid side merely removes the protecting soap film from the particles. The addn of Ca salts has much the same effect, by pptg the fatty acid as an insol soap. Other superficially active substances can bring about a reversible sepn of enzyme and substrate, e.g. ales in miscapillary proportions. The hormones, adrenaline, thyroxine and insulin, and the ergot poison, ergotamine, were practically without influence.

A. W. DOX

The problem of the specific action of tyrosinase. EMIL ABDERHALDEN and WALTER SCHAIKER. *Fermentforschung* 12, 329-32 (1931).—Previous expts with exts of dried champignon [*Agaricus?*] showed that the tyrosinase present was very reactive toward *p* tyrosine, adrenaline and 3,4-dihydroxyphenylalanine, but inert toward *o*- and *m*-tyrosine. The expts have now been repeated with practically the same results, except that a slight reddish coloration was noted at the surface of the *o* tyrosine and 2-amino-tyrosine solns. The same effects were produced by a tyrosinase prepn from dried potato peelings. On the other hand, exts of *Russula delica* gave a pos reaction with *m*-tyrosine, also with *o* and *m*-cresol. The reaction was neg with ephedrine, *d*- and *l*-ephedrine and thyroxine. The tyrosine dipeptides, glycyl-*l*-tyrosine, *d*- and *l*-alanyl-*l*-tyrosine, aminoisobutyryl-, *d* valyl-, *d*- and *l*-norleucyl-, *d*- and *l*-leucyl-*l*-tyrosine and the corresponding haloacyl derivs all reacted at pH 6.5 with formation of the characteristic red color. 2,5-Iodo- and 2,5-dibromotyrosine reacted, also *l*-tryptophan, while *d*-alanyl-*l*-tryptophan remained colorless. Possibly the pos result now obtained with *m* tyrosine is due to the presence of an addnl enzyme.

A. W. DOX

The active group of catalase II. KAAL ZEIRE *Z. physiol. Chem.* 195, 39-48 (1931) cf. *C. A.* 25, 119—The close relationship between porphyrin Fe content and catalase activity of horse liver ext. is further shown by the constancy of this ratio in preps purified by $\text{Ca}_3(\text{PO}_4)_2$ adsorption, dialysis, $\text{Al}(\text{OH})_3$ adsorption, and $\text{Ca}_3(\text{PO}_4)_2$ adsorption after 76% inactivation. The same relationship is now shown to hold for catalase preps from plant sources. Expts with the fungus, *Boletus scaber*, were unsatisfactory because of low enzyme activity and the presence of hemin in some other form which gave the hemochromogen reaction. A suitable source of plant catalase was found in the cotyledons of germinated pumpkin seeds. The max. activity develops between the 3rd and 5th days of germination. These cotyledons yield an ext. with $1/14$ the catalase activity of liver ext. The ratio of enzyme to porphyrin Fe as detd. by the hemochromogen spectrum is approx. 8000, about 3 times as high as that of liver ext. The dissoci. const. of the HCN compd. of pumpkin catalase is about $1/4$ that of liver catalase. In both enzymes the active group is probably of similar constitution. A. W. Dox

Activators Z II. A. FULPA AND TOBE PHILIPSON *Z. physiol. Chem.* 195, 81-100 (1931) Caffeine, theobromine, xanthine, tryptophan, adenosine, thymidine and hemin in concns. of 1-5 mg. per cc. have no influence on the rate of fermentation, even in the presence of an activator Z prep. Plant material from various sources was examd. for the presence of activator Z but the evidence was very slight except in malt germs where the ext. showed practically the same activity as boiled yeast ext. Z consists of 2 factors which are differentiated by their temp. stability and, like vitamin B, are more sensitive to alkali than to acid. Various pptn. and adsorption reactions were applied with a view to purifying the Z factors. A concd. yeast dialyzate was freed from certain impurities by treatment with $\text{Pb}(\text{OAc})_2$ and $\text{H}_2(\text{OAc})_2$. The filtrate was then used for adsorption expts. Treatment with $\text{AgNO}_3 + \text{Ba}(\text{OH})_2$ removes some of the activity. Picric acid does not ppt. the activator, since the loss of activity is not restored by recombining the ppt. and filtrate. Addn. of MeAc to 95% gives a ppt. which contains the full activity and about 40% of the inert material is left in the filtrate. $\text{Fe}(\text{OH})_3$ gives a ppt. from which the Z activator can be recovered by elution with 1% KH_2PO_4 but without the preliminary Pb - H_2 treatment the adsorption of activator does not occur. The Fe filtrate does not give up its active substance to charcoal regardless of the pH of the medium. A good adsorbent is $\text{Ba}_3(\text{PO}_4)_2$ when this is prepd. in the presence of the activator; otherwise no adsorption is effected. Recovery of the activator from this ppt. cannot be accomplished by elution but requires a decompn. of the phosphate by H_2SO_4 . Al^{+++} has the advantage of being insol. in AcOH and the activator which it adsorbs can be released by AcOH extn. and thus freed from phosphate. $\text{Al}(\text{OH})_3$ gives the same adsorption as $\text{Fe}(\text{OH})_3$. The activation of yeast by activator preps. reaches a max. beyond which further addn. is ineffectual. Bottom yeast is activated to the same extent by the original dialyzate as by the Fe filtrate; hence it is influenced only by that part of the Z activator which is not adsorbed by $\text{Fe}(\text{OH})_3$. A. W. Dox

The nature of the proteases VIII. Refractometric studies on the activity of various pepsin preparations I. A. SMOGONIN, A. N. ADOVA AND S. S. DAOSNOV *Z. physiol. Chem.* 195, 113-20 (1931), cf. *C. A.* 25, 1268—The sol. products resulting from digestion of an insol. protein, such as fibrin, increase the η of the medium. By detg. the η of the soln. after enzyme and substrate have reacted under standard conditions it is possible to compare the activity of various pepsin preps. Control expts. with enzyme alone show that the more active preps. have higher η values. A. W. Dox

Studies on the intermediary metabolism of tryptophan. I. Kynurenine, an intermediary metabolic product of tryptophan. Y. KOTAKE AND J. IWAO *Z. physiol. Chem.* 195, 129-47 (1931)—When tryptophan is injected subcutaneously into rabbits whose metabolism has been lowered by a regime of polished rice, a product intermediate between tryptophan and kynurenic acid is excreted in the urine. The name *kynurenine* is proposed for this substance. It the kynurenine is present in considerable amt. it can be crytd. out directly as sulfate by addn. of H_2SO_4 and a little Et_2O , but ordinarily its isolation requires a pptn. by phosphotungstic acid or H_2SO_4 and decompn. of the ppt. by $\text{Ba}(\text{OH})_2$ or H_2S . Kynurenine is $\alpha\text{-H}_2\text{NC}_4\text{H}_4\text{C}(\text{CO}_2\text{H})\text{CHCH}(\text{NH}_2)\text{CO}_2\text{H}$. It m. 190° (decompn.) with previous sintering and darkening, and has $[\alpha]_D^{25} -28.5^\circ$. It forms a sulfate which begins to darken at 165° and becomes black at 180° , and has $[\alpha]_D^{25} 10.7^\circ$. Esterification by satg. the EtOH soln. with HCl gives a diethyl ester di. HCl salt. The $\alpha\text{-NH}_2$ is shown by the ninhydrin reaction, and the double bond in

the side chain by the addn of 2 Br to form a *dibromide*, m 206–7° (decompr), which is the lactone of the aromatic NH_2 and the $\gamma\text{-CO}_2\text{H}$, and is thus an indole deriv. Excretion of kynurenine is accompanied by a decrease in kynurenine acid excretion. On the other hand, when kynurenine is fed to rabbits 20–40% may be recovered as kynurenine acid. II. The cleavage of kynurenine by means of barium hydroxide. Y KOTAKE AND M KRYOKAWA *Ibid* 147–52—Kynurenine is very stable to acids and withstands 10 hrs' boiling with 25% H_2SO_4 . It is very unstable to alkalis, splitting off NH_3 and CO_2 . After boiling with 1% $\text{Ba}(\text{OH})_2$, the products isolated were *o*-aminobenzophenone in 70–81% yield (identified as the HCl salt, m 264.5°, the Ac deriv m 75–6°, the mono-Br deriv of the latter, m 157–8°), and kynurenine acid. The reaction is best explained by the intermediate formation of *o*-aminobenzoylpyruvic acid which on the one hand would yield *o*- $\text{H}_2\text{NC}_6\text{H}_4\text{Ac}$ and $\text{C}_6\text{H}_5\text{O}$, by hydrolysis, and on the other hand kynurenine acid by enolization and ring closure to the quinoline structure. These products of hydrolysis furnish further proof of the structure of kynurenine (see above). III. The cleavage of kynurenine by means of sodium bicarbonate solution. Y KOTAKE AND G SCHICHIKI *Ibid* 152–8—When refluxed with 2% NaHCO_3 kynurenine splits off NH_3 and the soln becomes canary yellow, then orange yellow and soon shows fluorescence. Some *o*- $\text{H}_2\text{NC}_6\text{H}_4\text{Ac}$ is formed, but much less than in the cleavage by $\text{Ba}(\text{OH})_2$. After removal of this by Et_2O extn, the orange yellow product may be obtained from the mother liquor by acidifying and again extg with Et_2O . The substance crystallizes on evapn of the Et_2O but is unstable to air light and heat. Its m p varies considerably with different preps, hence the substance is probably not pure. Further treatment with NaHCO_3 or $\text{Ba}(\text{OH})_2$ yields *o*- $\text{H}_2\text{NC}_6\text{H}_4\text{Ac}$ but no NH_3 . When treated with warm water part of the substance dissolves and part forms an oil. A soln of the latter in Et_2O shows an intense blue fluorescence and yields yellow crystals of kynurenine yellow, m 182°, an unstable substance closely related to kynurenine acid. The orange yellow product obtained from kynurenine by NaHCO_3 cleavage probably consists mainly of *o*-aminobenzoylpyruvic acid, a part of which appears to be present in a condensation form. Kynurenine gives a blue color reaction when its aq soln is treated with PhNHNH_2 , HCl and NaHCO_3 , the mixt. heated to boiling, then cooled and treated with concd H_2SO_4 . IV. The mechanism of kynurenine acid formation in the organism. Y KOTAKE *Ibid* 158–66—The discovery of kynurenine as an intermediate in the metabolism of tryptophan to kynurenine acid, and the elaboration of its structural formula, furnishes an important link in the chain of reactions involved in the breakdown of tryptophan. The successive reactions are now believed to occur as follows: (1) tryptophan is oxidized to its Py α -hydroxy deriv, (2) the pyrrole ring is opened by hydrolysis to γ -*o*-aminobenzylglutaric acid, (3) removal of 2H from the glutamic acid grouping yields kynurenine, (4) CO_2 and NH_3 are removed by oxidation, yielding *o*-aminobenzoylpyruvic acid, (5) the tautomeric form of the latter closes to a quinoline ring with formation of kynurenine acid. V. The formation of kynurenine acid from kynurenine by microorganisms. G SCHICHIKI AND M KRYOKAWA *Ibid* 166–71—Synthetic culture media to which kynurenine had been added as the source of N were inoculated with *Osidium lactis* and *Willia anomala* and incubated for 17 and 33–89 days, resp. Both organisms converted part of the kynurenine into kynurenine acid, which was isolated and identified by m p and by analysis of the Ba salt. From the mold culture a small quantity of a substance m 122° was obtained by Et_2O extn; it was probably anthranilic acid. The yeast culture contained a colored substance possessing the general properties of urochromogen. A medium contg. α,γ -quinolinedicarboxylic acid gave a good growth of mold but no kynurenine acid. From kynurenine the max. yield of kynurenine acid was 60%. The cleavage of kynurenine by yeast is essentially the same as that by the animal organism, and supports the view that the formation of kynurenine acid and urochrome from tryptophan in the animal organism occurs via kynurenine. VI. The excretion of kynurenine acid in the bile and its stability in the organism. Y KOTAKE AND K. ICHIHARA *Ibid* 171–9—The path of excretion of the kynurenine acid formed in the organism from tryptophan varies with different animals. With dogs the excretion is often greater in the bile than in the urine, while with rabbits the excretion is entirely in the urine. This explains the variable results reported on dogs where only the urine was examd. Kynurenine acid, after subcutaneous injection as the Na salt, is excreted in both bile and urine by rabbits as well as dogs. It is not utilized by the organism, but is eliminated quantitatively, although only slowly after large doses. Kynurenine acid is not an intermediary product but an end product of tryptophan metabolism. VII. The site of formation of kynurenine acid in the organism. K. ICHIHARA, S. OTANI AND J. TSUJIMOTO (WITH COLLABORATION OF Y OKAGAWA AND T. KIYOMATSU). *Ibid*

179 84 Tryptophan was administered subcutaneously to dogs that had been hepatectomized after connecting the portal vein to the vena cava by means of a cannula. The animals lived 9.5 and 12.5 hrs., resp., at the end of which time the blood and urine were examined for kynurenic acid. The results were neg. A perfusion expt. in which defibrinated blood contg. kynurenic acid was passed 18 times through the isolated liver gave an 18% yield of kynurenic acid. In a control expt. without kynurenic acid no kynurenic acid was obtained. It is evident that conversion of kynurenic acid to kynurenic acid occurs in the liver. The excretion of kynurenic acid is increased by extirpation of the spleen. VIII. The question of the origin of urochrome. Y. KOTAKE AND H. SAKATA (WITH COLLABORATION OF Y. SHIBABE AND S. OTANI) *Ibid* 181-91.—After ingestion of 3 g. of tryptophan the urochrome content of the alloxyproteic acid fraction of the urine as detd. colorimetrically, nearly doubled. The Ba and Ca salts prepd. from this fraction were easily sol. in H₂O and gave the Pauly diazo and the PbS reactions, but neither the biuret nor the Ehrlich diazo reaction. The N content of the urochrome fraction varied with the intensity of the color. When kynurenic acid instead of tryptophan was administered the Ehrlich reaction became pos. The urine then had a greenish color which became yellow on standing in the air. The urochrome or urochrome is formed from tryptophan by way of kynurenic acid. Oral administration of tryptophan to a patient with pulmonary phthisis gave a marked increase in diazo and urochrome values of the urine. IX. The effect of tryptophan on experimental anemia and its relation to the spleen. Y. OKAGAWA AND M. TATSUI *Ibid* 192-202.—The recovery from exptl. anemia produced in dogs by injection of Ph. NITRIL, HCl is somewhat more rapid as detd. by erythrocyte count and hemoglobin content of the blood if the animals are given injections of tryptophan every 2 days. If the anemia is produced 25-45 days after removal of the spleen, recovery is less rapid and is not accelerated by the tryptophan treatment. One year after splenectomy the tryptophan treatment is somewhat more effective in curing exptl. anemia. The spleen appears to play some role in the utilization of tryptophan for purposes of blood building. The liver contains somewhat more tryptophan than the kidney. The amts. present undergo practically no change within 49 days after extirpation of the spleen, but after a longer period the tryptophan content of the liver increases. This compensation of spleen function also supports the view that tryptophan can serve as building material for hemoglobin formation. X. *dl*-Indolelactic acid and its use in nutrition. K. ICHIHARA AND N. IWAKURA *Ibid* 202-7.—The indolelactic acid used by Jackson (*C. A.* 21, 3070) with neg. results as a substitute for tryptophan in feeding expts. was not the *dl* form as J. supposed but the *l* form as indicated by m. p. The *dl*-acid prepd. by autoclaving the *l* acid with H₂O(II), until optically inactive, and tested by mixed m. p. with synthetic *dl* acid obtained by reduction of indolepyruvic acid, gives positive results with rabbits. The acid recovered from the urine was mainly the *l* form, showing that the *d* acid was actually utilized. Feeding expts. with rats showed that a loss in wt. occurred when the animals received a tryptophan free ration consisting of hydrolyzed casein, cystine, tyrosine, starch, sugar, salt and agar. When this ration was supplemented by *dl* indolelactic acid the animals more than made up the loss. Indolepyruvic acid and *l* indolelactic acid were ineffectual as substitutes for tryptophan. XI. The effect of methyltryptophan on artificial anemia and nutrition. Z. MATSUOKA AND T. NAKAO *Ibid* 208-14.—As measured by increase in erythrocyte count and hemoglobin content and decrease in reticulocytes, the recovery of guinea pigs from exptl. anemia is more rapid with α -methyltryptophan than with tryptophan administered subcutaneously. On the other hand the feeding of α -methyltryptophan to rats does not compensate a dietary deficiency in tryptophan. The action of methyltryptophan, and probably that of tryptophan, in curing exptl. anemia is independent of their nutritional value.

A. W. DOX

The chemical nature of urease. ERNST WALDSCHMIDT LEITZ AND FELIX STEIGERWALDT *Z. physiol. Chem.* 195, 260-6(1931).—The protein character of cryst. urease brings up the question whether the enzyme is actually a protein or whether the protein is merely a carrier which stabilizes the enzyme. The effect of activators such as amino acids and HCN is to overcome the inhibitory effect of accessory substances, since the activation is observed only with crude enzyme and not with purified or cryst. preps. The protein is not an essential component of urease, as shown by the fact that proteolysis by pancreatic ext., trypsin or papain HCN to the point where sulfosalicylic acid no longer gives a ppt. does not diminish the activity of the enzyme. Hydrolytic products resulting from the breakdown of this protein function equally well as carriers of the urease activity. The protein may even be only an accessory concomitant of the pure enzyme.

A. W. DOX

The intermediary metabolism of histidine. III. S. EDLBRACHER AND J. KRAUS. *Z. physiol. Chem.* 195, 267-72 (1931), cf. *C. A.* 25, 118.—In the enzymic cleavage of histidine, previously reported, the products isolated were glutamic acid, HCO_2H and 2NH_3 . The method of isolation, however, involved the use of NaOH and HCl , and the possibility remains that the end products were thus formed from an unknown intermediate. Kauffmann and Mi-lowitzer (*C. A.* 25, 307) have since shown that the use of Na_2CO_3 instead of NaOH yields only NH_3 . This observation is now confirmed by E. and K. The possible intermediates glutamine and HCO^-NH_2 show the same difference in behavior toward Na_2CO_3 and NaOH . Both differ, however, from the histidine intermediate in that the former is hydrolyzed by dil. HCl while the latter is not. In the progress of the enzymic cleavage of histidine the ratio of NH_3 liberated by Na_2CO_3 to that liberated by NaOH decreases from 0.83 after 2 hrs. to 0.53 after 24 hrs., where it then remains const. The first attack of the enzyme breaks open the imidazole ring and liberates NH_3 . The product is probably an amide which is hydrolyzed by NaOH but is resistant to HCl . Whether a further breakdown is due to enzymic action or to a spontaneous decomposition has not as yet been detd. A. W. DOX.

An easily cleavable phosphoric acid compound in yeast. HANS v. EULER AND RALFAR NILSSON. *Z. physiol. Chem.* 195, 273-8 (1931).—Ext. of dried yeast, after removal of protein by $\text{CCl}_4/\text{CS}_2\text{H}_5$, contains a substance from which H_3PO_4 is not pptd. by magnesia mixt. and NH_4OH but which gradually forms a ppt. with Embley's reagent. The H_3PO_4 is thus easily split off. The amt. of this complex rapidly diminishes during the first few hrs., but less rapidly in the presence of added $\text{Na hexoediphosphate}$. In the presence of H_3PO_4 it increases while the H_3PO_4 disappears within 2 hrs. then the amt. rapidly decreases. Addn. of creatin retards this synthesis. A. W. DOX.

Pepsin. Preliminary paper. HEINZ HILFERT. *Z. physiol. Chem.* 195, 1-9 (1931).—The early stages of peptic digestion were studied by measuring the changes in viscosity of a 3% casein soln. at pH 2.3 and 30° . Dil. pepsin solns. were used in order to prolong the initial stage and make possible a greater no. of observations. The viscosity, as measured by timing the flow through a viscometer, decreases in the very first phase of the digestion and soon reaches a min., after which it rapidly increases and attains a sharp max., which is a multiple of the original value. This stage marks the coagulation of the entire casein soln. As the digestion proceeds there occurs a sepn. of phosphopeptone, and with increasing d. of the floccules the viscosity again decreases. The peptic cleavage was followed by means of the Linderström-Lang acid titration in acetone. The initial digestion stage in which viscosity decreases was studied in particular. Viscosity change is not a simple function of cleavage but varies with the nature of the peptide linkages broken. The ratio between cleavage and viscosity is different for different pepsin preps. but const. for one and the same prepn. The curve can serve for characterization of a pepsin digestion just as the location of pH optimum toward definite substrates distinguishes different proteolytic enzymes. Studies were made of different pepsin preps., including Northrop's crysd. pepsin and fractions obtained by adsorption on $\text{Al}(\text{OH})_3$. The sepn. fractions give different curves, but when they are recombined the original curve is again obtained. Even the crysd. pepsin is amenable to fractionation. Peptic action is therefore not a simple phenomenon, and the variations noted must be due to differences in the enzyme itself. A. W. DOX.

The physiological actions of cozymase. SVEN GARD. *Z. physiol. Chem.* 195, 65-70 (1931).—Injection of cozymase into the jugular vein of rabbits under urethane narcosis causes a lowering of blood pressure. This effect is independent of the specific action of cozymase, since inactivation by heating does not diminish it. Pure adenylic acid of yeast is somewhat more active in this respect, probably because of the fact that cozymase is not an entirely pure adenylic acid. Adenylthiomethylpentose has about the same effect. Introduction of the thiomethyl group into the adenosine mol. apparently does not diminish this special action. In the surviving rabbit heart, on the other hand, cozymase injection is followed by a considerable increase in activity, perhaps on account of dilation of the coronary vessels. Here also there is no difference in effect between active and inactive cozymase. The action of cozymase in lowering the blood pressure is a non-specific effect of adenylic acid or homologous adenosine derivs., which corresponds to that observed by Zipf with muscle adenylic acid. A. W. DOX.

Cytochrome. R. BIRNICH AND A. ROSENBAUM. *Z. physiol. Chem.* 195, 87-9 (1931).—The discrepancy pointed out by Shibata and Tamiya (*C. A.* 24, 3532) between his results and those of B. and R. (*C. A.* 24, 177) with regard to the influence of KCN in preventing the disappearance of the cytochrome band by oxidation is due to the fact that the cytochrome in the first instance was of vegetable origin (yeast) while in

the 2nd it was of animal origin (rat testes). In animal tissues the B component is oxidized in the presence of KCN while the C and A components remain reduced.

A. W. DOX

The absorption of ultra-violet light by amino acids (fragments of proteins). H. GUTTMANN, K. SCHWIERIN AND F. STÄHLER. *Strahlentherapie* 39, 401-41 (1931). — In synthetic mixts. of the amino acids corresponding in compn. to albumin, globulin and fibrin 2 absorption maxima are present, at 300μ and 250μ , which are the specific bands for the action of the ultra violet rays on the skin. Tyrosine, phenylalanine and glutamic acid are the substances which absorb the ultra violet rays very powerfully. In substances of the propionic acid series (propionic acid, alanine, phenylalanine, tyrosine, histidine and tryptophan) the influence of various groups on the absorption power is entirely independent of the mol. wt., but depends on the stereochem. structure of the mol. The most powerful in this respect is the aldehyde group, the activity decreases in the order NH_2 , CO_2H , OH .

F. R. GREENBAUM

The biologic action of Röntgen rays. Action of Röntgen rays on the potassium and calcium in the blood serum of man. G. SCHÄAL, M. A. GRUSCHETSKAYA AND E. J. ZWILICHOWSKAY. *Strahlentherapie* 40, 111-23 (1931). — Röntgen rays caused a change in the quotient K/Ca in the human blood serum. The direction of this change seems to depend upon the time interval between irradiation and blood examn. After the first day of the radiation the K seems to be increased but later on the Ca seems to be increased. This may account for the following phenomena: fluctuation of a no. of leucocytes, analgesic action, early reactions and poisoning due to Röntgen treatment.

F. R. GREENBAUM

Biologic investigations on the transparency of gypsum and mica to ultra-violet. TOMOJUNE TAKAGI. *Strahlentherapie* 40, 189-92 (1931). — Light from a Hg arc lamp which has passed through natural plates of gypsum is very active on blood agar plates as well as on the human skin; light filtered through mica was without action.

F. R. GREENBAUM

Effect of hydrogen-ion concentration upon the precipitation of certain basic substances by phosphotungstic acid. RUDOLF A. PETERS. *Biochem. J.* 24, 1852-5 (1930). — With most bases the pptn. appears upon passing from a more alk. to a less alk. soln. with guanidine compds. this is reversed.

BENJAMIN HARROW

Inhibition of esterases by excess substrate. DAVID R. P. MURRAY. *Biochem. J.* 24, 1899-6 (1930). — Haldane's suggestion ("Enzymes," p. 84 (1930); cf. C. A. 25, 720), that the inhibition of certain enzymes by excess of substrate is due to the formation of a non hydrolyzable complex of the enzyme with 2 mols. of the substrate, has been tested for the esterase of sheep liver and found to explain the facts.

BENJAMIN HARROW

Crystalline urease: a review. HENRY TAUBER. *J. Am. Inst. Homeopathy* 24, 16-8 (1931). — The literature on urease is reviewed. A bibliography is given.

JOSEPH S. HEPBURN

Prothrombase: its preparation and properties. JOHN MELLANBY. *Proc. Roy. Soc. (London)* B107, 271-85 (1930). — Prothrombase is prepd. from oxalated mammalian serum by (1) pptn. of its globulin complex at pH 5.3 after dial. with 10 vols. of H_2O , (2) extn. from this complex with a dil. soln. of $\text{Ca}(\text{HCO}_3)_2$, and (3) pptn. from the resulting soln. by addn. of acid to pH 5.3. Approx. 40 mg. prothrombase results from 100 cc. plasma, as a rule, 1 mg. of the prepn. after its conversion into thrombase, coagulates 100 cc. plasma in 20 sec. Prothrombase is a white amorphous powder giving the characteristic reactions of acid metaprotein; it is insol. in H_2O , readily sol. in dil. alkali and not dialyzable through a cellophane membrane. Heating at a temp. of 100° for 5 min. does not destroy prothrombase, but diminishes the rate at which it may be activated. Activation is produced by addn. of thrombokinase, is initially slow and proceeds with a rapidly increasing velocity, is influenced by the reaction of the medium and the temp., is accelerated greatly by the presence of small quantities of Ca or Sr and slightly by Ba or Mg.

JOSEPH S. HEPBURN

Porphyrin of component c of cytochrome and its relationship to other porphyrins. R. HILL AND D. KEILIN. *Proc. Roy. Soc. (London)* B107, 286-92 (1930). — Unmodified cytochrome c of bakers' yeast yields 2 different porphyrins. The one porphyrin, obtained by the action of HBr and AcOH, is indistinguishable from hematoporphyrin having the following properties of the latter: (1) soly. in Et_2O and AcOH, (2) absorption spectra in Et_2O and in HCl, (3) combination with Fe, various compds. of N, and globin to form hemochromogens or hemoglobins, (4) conversion into protoporphyrin by (a) heating *in vacuo*, or (b) heating the corresponding hematohematin *in vacuo*, or (c) heating the hematohematin in AcOH in a sealed tube. The other porphyrin results from the action of HCl in the presence of a reducing agent such as SO_2 ; it is insol. in

dextrin in which the I reaction was yellow, and kept for 10 days at 20° and pH 6.2. The soln gives a violet blue color reaction for I IK. From the soln the synthesized dextrin was pptd. It was proved by the reactions, the soly., the viscosity and mol wt. to be a polymerized dextrin. The existence of amylocoagulase in the malt ext could not be verified. But coagulation of starch by the yeast ext was found at any concn. Amylocoagulase and amylomutase exist in the yeast ext. Sol starch was hydrolyzed as much as possible by malt amylase and then the yeast ext. was added. The I reaction became violet after 18 days. The synthesized dextrin was pptd. by alc. The hydrolysis of the synthesized dextrin by amylase was like that of starch.

Y. KIHARA

Proteins of wool. II. MICHIO SAITO *J Agr Chem Soc Japan* 6, 1109-24 (1930).—The wool was treated with 1% HCl under 5 atm pressure for 5 hrs. Ninety % of the solid matter was dissolved. By fractional pptn 6 albumoses and a pepton were obtained. These are compared.

Y. KIHARA

Glucuronic acid fermentation. TEIZO TAKAHASHI AND TOSHINORI ASAI *J Agr Chem Soc Japan* 7, 1-5 (1931).—By the fermentative action of *Bact induratum* var. Hoshigaki on a medium consisting of 11 yeast ext., 100 g glucose and 25 g CaCO₃ for 30 days 35 g Ca hydroxygluconate was obtained. From the filtrate 20 g Ca gluconate was obtained. By the addn of MeOH to the filtrate of Ca gluconate, 20 g, Ca glucuronate was pptd. Ca glucuronate was proved by the color reactions, sp. rotation of the free acid ($[\alpha]_D^{25} = -18.5^\circ$) and Ca salt ($[\alpha]_D^{25} = -24.0^\circ$), phenyl osazone, m 131-2°, p-nitrophenylhydrazone, m 209-10°, Ba salt of p-bromophenyl osazone, m 204-7°, and brucine salt, m 178-9°. From the fermentation products of *Bact induratum* var. Hoshigaki on a medium of yeast ext. and Ca gluconate, 5 g Ca glucuronate was obtained. The bacteria produce glucuronic acid and hydroxygluconic acid not only from glucose, but also from gluconic acid.

Y. KIHARA

Amylase from wheat. DATTATREYA V. KARMARKAR AND VINAYAK N. PATWARDHAN *J Indian Inst Sci* 13A, 159-64 (1930).—The amylase from wheat is more active than that from barley. Its optimum temp is 49-58°, pH 4.6. Heating the dry enzyme at 160° for 1 hr destroys its activity, which is not regenerated by adding salts or amino acids. Inactivation results on prolonged dialysis through collodion bags or in the presence of pepsin. Neutral salts and amino acids are without influence.

MARY E. LEAR

Isolation of protein crystals possessing tryptic activity. JOHN H. NORTHROP AND M. KUNITZ *Science* 73, 262-3 (1931).—A cryst. protein has been isolated from a com. prepn. of "trypsin" which digests casein and gelatin in neutral solns. The digestive power of the crystals is about 10 times that of the most active com. preps. The activity remains const. through 3 successive crystns. The substance is unstable and may become less active during its prepn. unless proper care is exercised. The process of prepn. of the substance is briefly outlined.

B. S. LEVINE

The action of blood on acetylcholine. K. MARTIN *J. Physiol* 70, 333-48 (1930).—The destruction of acetylcholine by blood is due to the action of an esterase, which is specifically inhibited by eserine and also by fluorides.

J. F. LYMAN

The inactivation of histamine. C. H. BEST AND E. W. McHENRY *J. Physiol.* 70, 349-72 (1930).—The destructive agent, which is in highest concn. among the tissues of the dog in kidney and intestine, appears to have the properties of an enzyme; the name histaminase is suggested. Histamine seems to be destroyed by it through a change involving rupture of the imidazole ring.

J. F. LYMAN

A physicochemical theory of the action of substances on living matter. P. LAZAREV *J chim phys* 28, 42-8 (1931).—It is assumed that the functions of a cell depend on the concn. of the material introduced C_0 , and that the sensibility E can be expressed in such measurable quantities as rate of oxygen consumption, speed of nervous reaction, etc. The equation $E = E_0 + E_1 \beta C_0 (1 - e^{-\alpha t})$ is derived to show the sensibility of the cell at time t , where β is the ratio of the diffusion rate into the cell to the rate out of the cell, the latter being α . The equation is applied to the action of Et₂O on assimilation by chlorophyll and to the action on the nervous centers of substances introduced into the blood.

J. G. McNALLY

The action of monochromatic light. GUIDO GUERRENT *Boll soc ital biol sper* 5, 1098-1100 (1930) cf *C A* 25, 534.—The action of light filtered through colored plates on *S. cerevisiae* in the presence of glucose was studied. The amt of CO₂ liberated as a function of time was max. for light filtered through red plates; slightly less for light filtered through yellow plates, still less for light filtered through green plates, and nothing or almost nothing for light filtered through blue plates. Light rays of long

wave length therefore stimulate the action of *S. cerevisiae* in the presence of glucose

PETER MASUCCI

Alcoholic fermentation in liver. VIRGILIO MARTINI *Boll. soc. ital. biol. sper.* 5, 1172-5(1930)—Liver pulp, strongly acidified with HCl (0.6-0.7%) and incubated at 37°, undergoes alcoholic fermentation with the formation of CO₂ and EtOH. The fermentation is caused by some thermolabile factor

PETER MASUCCI

Precipitation of animal proteins. GEORG GRASSER *J. Faculty Agr. Hokkaido Imp. Univ.* 27, Pt. II, 227-42(1930)—The behavior of solns of gelatin and egg albumin with a large no. of phenols, naphthols, aldehydes and inorg. salts is tabulated.

H. B. MERRILL

Electrokinetic phenomena. III. The "isoelectric point" of normal and sensitized mammalian erythrocytes. HAROLD A. ABRAMSON *J. Gen. Physiol.* 14, 163-77(1930), cf. *C. A.* 24, 5568—Published isoelec. points of certain mammalian erythrocytes derived from studies of electrophoretic mobilities are the result of equil. incidental to erythrocyte destruction. This is a study of the mobilities of washed sheep and human erythrocytes in 0.85% NaCl soln at pH 3.6 to 7.4 made in 2 min. after prepn. of the suspensions. Under these conditions reversal of the sign of elec. charge did not occur, it occurred only after time had elapsed for the adsorption of the products of red cell destruction. Mobility changes little with decrease of pH. The reversal of the sign of the charge occurs in the presence of normal and immune (anti-sheep) rabbit sera. The isoelec. point detd. under these conditions is apparently not specifically connected with the immune body but may be incidental to red cell destruction and the presence of serum. The characteristic lowering of mobility by amboceptor, however, occurs from pH 4.0 to 7.4. The identity of the isoelec. points for normal and sensitized sheep cells is apparently not concerned with the immune reaction. Red cells are probably not covered with a film of albumin or globulin. Serum protein reacts with red cells in acid soln whereas gelatin does not. Lowering of mobility, usually ascribed to anti-sheep rabbit serum, may also occur, but to lesser degree, in normal rabbit serum. This diminution of mobility is not assoc. with sensitization to hemolysis induced by complement. Only a very small part of the surface of the red cell need be changed to obtain complete hemolysis in the presence of complement.

C. H. RICHARDSON

The kinetics of penetration. III. Equations for the exchange of ions. W. J. V. OSTERHOUT *J. Gen. Physiol.* 14, 277-84(1930); cf. *C. A.* 24, 5318—This is a preliminary attempt to outline the math. treatment of an exchange of ions between the interior of a living cell and the external soln.

C. H. RICHARDSON

The reversion of protein hydrolysis. A. SARLUVY *Arch. néerland. physiol.* 16, 136-44(1931)—The idea of Wasteneys and Borsook that such simple aromatic compds. as benzene, toluene, xylene, benzoic acid or benzaldehyde when added to peptic protein hydrolyzates synthesize proteins is not held by S. It is his belief that the addn. of BzH to these hydrolyzates forms Schiff's base and that the addn. of benzene, toluene and xylene is merely an adsorption phenomenon.

M. H. SOULE

The action of muscle phosphatase, cozymase and insulin. J. BODNÁR AND B. TANKÓ *Magyar Orvosi Arch.* 31, 105-31(1930); cf. *C. A.* 25, 1889—The amt. of esterified phosphoric acid increases remarkably if the phosphate concn. is increased, 0.25 g. muscle powder (prepd. by Euler's method, from pigeon breast) esterified in max. 11.42 mg. P₂O₅ (4.57%). The original muscle powder is not affected by boiled muscle juice. Muscle powder washed with H₂O at room temp. inhibits considerable of its activity, which is due to the want of cozymase, since the boiled juice increases the phosphorylation of the washed powder in the presence of higher phosphate concn. Washed muscle powder esterified 0.63%, washed muscle-powder plus boiled muscle juice 2.03% P₂O₅. The boiled muscle juice not only increases the esterification of the washed muscle powder but it has the power of counteracting the lowering of the optimum caused by the injury of the phosphatase. Boiled yeast juice also has the power of re-activating the esterification of the washed muscle powder. Fresh muscle paste loses its activity in a few hours, but the dry powder keeps rather long. In 5 weeks only 20% of its activity is lost. Insulin has no effect on the power of phosphorylation either in fresh muscle paste or in muscle powder.

H. TAUBER

Thermal disaggregation of gelatin (GREGGROSS, *et al.*) 2. Chemical decomposition by radiation (HARKEE) 3. A new physicochemical explanation of the formation of humus, peat and coal. The significance of biological factors in these processes (ZOLCINSKI) 8. Some relations between constitution and odor (ANGELI) 10.

DOGNON, ANDRÉ: Compendio de fisico-química biológica y médica. Translated from the French by Germán García. Madrid: Espasa-Calpe 318 pp.

NLEDHAM J. SLETH A Chart to Illustrate the History of Biochemistry and Physiology Cambridge, Eng the author 2 charts Paper, 10s. 6d Reviewed in *J Am Med Assoc* 96, 1100(1931)

Vegetable albumins HENRI A BEAUFORT Fr 695,233, Aug 30, 1929 Col-
loidal solns of the albumins are treated in a salt medium by dialysis and ultrafiltration
in 2 successive phases The 1st phase consists in the use of the phenomenon of dialysis
and osmosis or electro-dialysis and electro-osmosis an ultrafilter membrane of medium
porosity is used as dialyzing membrane Between the 2 phases a large part of the
albumin is sepd by hypercentrifugation this process depends on the change of mol
state provoked by the dialysis Cf C A 24, 5051

Enzymes ALBERT R BORDIN and IVAN A FERRONT Fr 37,420, May 23,
1929 Addn to 670,099 (C A 24, 2007) The pH of worts fermented to obtain
proteolytic enzymes of high power is corrected during this fermentation as often as
necessary by the addn of compds of alk, acid or amphoteric reaction

Enzymes I G FARMENING A G Ger 515,931, Feb 18, 1928 Enzymes
are prepd by treating the starting material such as milk, latex, etc, with proteases of
 pH about 7, sepg the coagulation product and isolating the enzymes therein in the
usual manner Cf C A 24, 4196

B—METHODS AND APPARATUS

STANLEY A BENEDICT

A simplified method for quantitative tissue culture in vitro. MACHTELD SANO
AND LAWRENCE W SMITH *Proc Soc Exptl Biol Med* 28, 282-4(1930)—The tissue
is grown in media contained in a special chamber consisting of a Pyrex glass ring 3 in
in diam, 5-6 mm high with parallel flat ground surfaces, 2-3 mm in thickness. Thin
sheets of mica $3\frac{1}{4}$ in. in diam form the top and bottom of the chamber, they are
held in place by rims of petrolatum The chamber is kept in a 4 in Petri dish The
method permits the examn of the growing tissue with high power microscopic objectives
quant growth is possible and the tissues may be fixed and stained without distortion
C V BAILEY

Micro method for the determination of cystine in proteins. A D MARENZ
Compt rend soc biol 104, 405-7(1930)—See C A 24, 5049 E J C

The methylene-blue staining of yeast cells and studies on the permeability of the
yeast-cell membrane I HERMANN FINE *Z physiol Chem* 195, 215-40(1931),
cf C A 24, 43-4—A crit study was made of the methylene-blue staining method
commonly employed for distinguishing between living and dead yeast cells, and also
a study of the permeability of the cell membrane Both the quantity of dye taken
up by the yeast and the no. of cells stained depend on the pH For one and the same
yeast quite different values may be obtained for the no. of "dead" cells according to
whether the yeast is taken directly from the fermented substrate or given a subsequent
treatment with tap water The influence of salts on the methylene-blue staining is
still more important In distd water the no. of stainable cells rapidly increases while
in tap water it remains practically const during the same period of time This differ-
ence becomes enormous when sugar is also present In a few min as many as 98%
of the cells in the salt free sugar soln are stained deep blue, while in tap water the
presence of sugar has little influence This result is not to be interpreted in the sense
of the methylene blue method as the death of the cells, but must be considered from the
standpoint of a change in permeability of the cell membrane By the brief sojourn in
an electrolyte free sugar soln the cells or their membranes are so altered that the
methylene blue can penetrate These cells thus prepd for the reception of dye and
easily stained deep blue are sound and capable of proliferation They are not killed
until the dye has actually penetrated The predisposition to reception of dye can be
completely removed again by a small addn of salt The same phenomena are evoked
equally by glucose, fructose, sucrose and galactose maltose is somewhat less effective
They are more pronounced with brewer's yeast than with baker's yeast The results
were confirmed and extended in eudiometer expts In electrolyte-free mixts the fer-
mentation velocity was less than in the presence of salts If methylene blue was also
present in the system the fermentation was strongly inhibited This inhibition is
almost completely overcome by salts, the activity of their ions following essentially the
order of the Hofmeister series

The macrodetection of alkaloids G D LANDER, *Am J. Pharm* 103, 100-3
(1931)—See C A 24, 5048 A W. DOX
W. G GAESSLER

Studies on the specificity of Asher's method for testing thyroid gland function by oxygen want. LEON ASHER AND HANS WAGNER. *Z ges expl Med* 68, 32-81(1929)
F. L. DUNN

The chemical detection of adrenaline in the blood. G VIALA. *Boll. soc. ital. biol sper* 5, 1163-7(1930).—To 1 cc of blood add 1 cc of satd HgCl_2 acidified with traces of HCl ($\text{pH} = 6$). Filter, add to the filtrate a few drops of satd. sulfanilic acid and satd KIO_4 soln., boil. The soln. develops a rose color. The reaction is always pos with blood, occasionally with lymph neg with cerebrospinal fluid. The reaction is not given by urea, uric acid, tyrosine, tryptophan, indole, cholesterol, ephedrine, nor by the products of acid hydrolysis of egg white. Animal expts indicate that the reaction is given by adrenaline. Since the blood of dogs deprived of the suprarenal capsules also gave a pos reaction, the hypothesis is advanced that the reaction is also given by precursors of adrenaline, substances contg the pyrocatechin group. The reaction may be used as an index of the amt. of adrenaline in the blood, by estg the quantity colorimetrically. It is sensitive in a diln. of 1 to 30 million.

PETER MASUCCI

Preferred methods for the clinical determination of glucemia. LUIGI PINELLI AND ATTILIO CORBIA. *Biochim terap sper* 18, 19-33(1931).—The methods of Folin-Thivolle, Hagedorn, Hartmann-Schaeffer, Ionesco-Matiu and Bertrand-Silvestri are described and the results compared. Generally the methods based on the reduction of Fe salts give higher results than the Cu methods. The av figures for normal persons by the method of Folin are 0.085, by Hagedorn 0.133, Ionesco 0.162, Schaeffer 0.076 and Bertrand 0.085%. The av in diseases was 0.088, 0.103, 0.157, 0.083 and 0.094, resp. The Ionesco method gave in some cases results 4 times as high as those by the Folin method.

A E MEYER

Estimation of Fe in biological material (HILL) 7.

OSGOOD, ERWIN E., AND HASKINS, HOWARD D. Laboratory Diagnosis. Philadelphia P Blakiston's Son & Co \$5

Determination of albumin according to Esbach. J VASARHELYI. *Hun.* 101,531, Dec. 26, 1929. Esbach tubes that can be used in centrifuges are specified. They reduce the time required for the detn. to a few min. instead of 24 hrs.

C—BACTERIOLOGY

CHARLES B MORRELL

Microbic dissociation with reference to the attenuated tubercle bacillus of Calmette and Guérin (B. C. G.). R S BECHIE. *Edinburgh Med J* 38, 174-82(1931).—Evidence is presented for the disson of B. C. G. into intermediate, rough and smooth types. By guinea pig inoculation the virulence is found to increase in the order named. A fresh culture of B. C. G. on glycerol bile potato medium is less virulent than a subculture on Dorset's egg medium.

RACHEL BROWN

Experimental and applied studies in electrosterilization. I. LOUIS I GROSSMAN AND J L T APPLETON, JR. *Dental Cosmos* 73, 147-60(1931).—With oral bacteria and a clinically tolerated dosage of current (30 ma-min), greater antibacterial effect was uniformly obtained from an electrolyzed soln. than from the same soln. without electrolysis. The antibacterial effect always was greater at the anode. ZnI_2 was more efficient than KI. Of approx. 50 solns. tested, the greatest antibacterial effect compatible with tissue tolerance was given by 1 M ZnI_2 in 0.1 M I; this soln. had no greater irritant action than 0.85% NaCl soln. The action exerted was bactericidal and due directly to the action of the electrolyzed compd., and not to a change in the reaction of the medium. II. Relative antibacterial effects of different electrodes used in electrosterilization. *Ibid* 230-3.—The antibacterial effect depended upon both the electrode and the electrolyte. With NaCl as the electrolyte, a Zn electrode gave a greater antibacterial effect than an Ir-Pt electrode. With either ZnCl_2 or $\text{ZnI}_2 + \text{I}_2$, both types of electrode gave similar results. As electrode material Wipal steel was more efficient than Pt, less efficient than Zn. Both Zn and Zn-coated electrodes were slightly corroded by 1 M ZnI_2 in 0.1 M I, while an Ir-Pt electrode was unaffected. III. Single- and multiple-electrode studies. *Ibid* 370-3.—The antibacterial effect from a single electrode is greater than that from each of the 2 or 3 electrodes constituting a multiple electrode. Each constituent of a multiple electrode has practically the same effect.

JOSEPH S. HEPBURN

Theories of the bacteriophage. JULES BORDET. *Proc Roy Soc. (London)* B107, 398-417(1931).—The bacteriophage is not a virus, but its intense action represents

merely the pathol. exaggeration of a normal function connected with mutation, the lysis being a transmissible autolysis.

JOSEPH S. HIPPBURN
Modification of the Morae-Kopeloff culture of anaerobes. ARAO ITANO AND SACHIO ARAKAWA. *J. Agr. Chem. Soc. Japan* 6, 976-8 (1930).—A mixt. of 0.5-1.0 g. pyrogallol and 10-15 cc. 10% KOH was put in a Petri dish, and covered by another dish with agar culture medium. The dishes were sealed with Toma tape. This is a convenient method for the culture of *Clostridium pasteurianum* and of thermobacteria at 60-5°.

Y. KIHARA
The influence of soaps on the germicidal properties of certain mercurial compounds. BETTYLEE HAMFILL. *Am. J. Hyg.* 13, 623-37 (1931).—The presence of Na oleate or a mixt. of soaps such as Lux in concn. of 1% or over greatly increases the toxicity of HgCl₂ to virulent staphylococci at 37°. Smaller quantities of soap gave varying results. Large amts. were no better than 1%. Destruction of the bacteria occurred in 5 min. The fact that there is no toxicity is not due to increased absorption due to the soap influencing the membrane, for it makes no difference whether the soap and the mercurial are used together or in any order. Some other organisms were not affected by the soap and for some its presence was protective. Soap increases the action of HgBr₂ and Hg(NO₃)₂ in some instances, particularly for *Staph. aureus*. The toxicity of Hg(CN)₂ and o-chloromercuriphenol is increased 3 to 4 times by soap.

G. H. W. JONES
Germicidal efficiency of soaps and of mixtures of soaps with sodium hydroxide or with paeonols. J. M. SCHAEFER AND F. W. TILLEY. *J. Agr. Research* 41, 737-47 (1930).—Coconut-oil soap is the only common soap with significant germicidal activity at room temp. Its activity against most of the organisms tested is comparatively low but in the absence of org. matter it manifests a very high germicidal activity against *Pasteurella avicida* and *P. ruisellica* and some strains of hemolytic streptococci. In the presence of org. matter, however, its activity toward these organisms is relatively low. The addn. of NaOH to solns. of neutral coconut-oil soap increases the effectiveness of the solns. against the various bacteria tested. Coconut-oil soap, castor-oil soap or linseed-oil soap when mixed with various phenolic compds. in the proportion of 1 part of soap to 2 parts of phenol considerably increased the germicidal efficiency of the phenols in the absence of org. matter. In the presence of milk or blood serum such increases were comparatively small. Germicidal soaps made by adding cresol, com. cresylic acid or orthophenylphenol to coconut-oil soap are efficient against *Staphylococcus aureus* even in the presence of milk or blood serum. There is no satisfactory explanation of the germicidal activity of soaps, or of their action in enhancing the germicidal activity of phenols.

W. H. ROSS
A System of Bacteriology in Relation to Medicine. Vol. VII. Virus Diseases and Bacteriophage. Issued by Medical Research Council at London. New York (551 Fifth Ave.) The British Library of Information. 509 pp. \$6. Reviewed in *Am. J. Pub. Health* 21, 462, *Nature* 127, 265 (1931).

D'HERELLE, F. The Bacteriophage and Its Clinical Applications. Translated by George H. Smith. London. Baillière, Tindall & Cox. 254 pp. 18s. Reviewed in *J. State Med.* 39, 248 (1931).

D-BOTANY

THOMAS G. PHILLIPS

Changes in the phosphorus content of growing mung beans. JAMES E. WEBSTER AND CLAUDE DALBOM. *J. Agr. Research* 41, 819-24 (1930).—A study of the P content of the various aerial organs of the mung bean as the plant grew to maturity shows that the stems and branches have a greater proportion of total inorg. P than is found in the leaves. The decrease in total P is much larger than the decrease in ash content. There is a relatively small change in the amt. of inorg. P in any 1 organ as the plant matures. There are indications that the P taken into the plant is metabolized to org. forms for storage. Lipoid P constitutes only a small fraction of the green plant and changes little during growth. There is only a slight change in the total P content at harvest times, so that differences of 2 or 3 weeks would make only negligible changes in the P feeding value. There is no evidence that P is transferred from the leaves and branches to the stems, and it is assumed that the localization of P follows the normal ash intake and is not the result of metabolic activities.

J. R. ADAMS
The chemical composition of *Ambrosia trifida* at successive growth stages. R. B. DUSTMAN AND L. C. SHRIVER. *J. Am. Soc. Agron.* 23, 190-4 (1931).—The crude protein, N free ext. and ash content of *Ambrosia trifida* are high well up to the time of

bloom while the fiber is correspondingly low. This plant might be made to yield considerable amts. of forage of high nutritive value because of its leafy nature and the high N content of the leaves.

J. R. ADAMS

Guttation and its relationship to nitrates and other salts in the nutrient medium. A DRAGHETTI *Ann. staz. sper. agr. Modena* [N. S.] 1, 123-32 (1930).—Wheat plants were cultivated on nutrient media containing several salts, and the quantity of liquid obtained by guttation was measured. This quantity always reaches a max. in the 4th or 5th day of culture. When the nutrient soln. contains Ca, Na, NH_4 , or K nitrates, the guttated liquid gives the nitrate reaction, but when Ba or Sr nitrates are present, the NO_3 reaction is not obtained.

G. A. BRAVO

Morphological, biometrical and chemical notes on the cherries cultivated in the Modena Province. G. D'IPPOLITO AND E. ALBERTI *Ann. staz. sper. agr. Modena* [N. S.] 1, 189-202 (1930).—The chem. analyses of the cherries made by Berard, Fresenius and Richardson are given (without bibliography) and the morphological characters of 11 varieties of cherries (derived from *Cerasus cerasus* L.) are described.

G. A. B.

Effect of turbulence (of the atmosphere) on the carbon-dioxide exchange in plants. P. LEHMANN *Fortschritte Landw.* 4, 745 (1929).—Variations in the CO_2 content of the air at short intervals and under varying conditions are recorded. The effect of wind velocity, temp. and environmental conditions on local changes in the moisture and CO_2 contents of air and the resulting influence on the CO_2 assimilation of plants, are discussed.

B. C. A.

Cyanogenetic glucosides in Australian plants. II. (a) *Eremophila maculata*. HORACE FINNEMORE AND CHARLES BERTRAM COX *J. Proc. Roy. Soc. N. S. Wales* 63, 172-8 (1930).—Prunasin (cyanogenetic glucoside) has been isolated from the bark of *Prunus serotina*, and its identity with the glucoside now isolated from *Eremophila maculata* has been confirmed by the mixed m.p. method. (b) Presence of enzymes in fodder plants as a factor in the poisoning of stock. SUZANNE K. REICHERD. *Ibid.* 179-82 (1930).—The isolation of sambunigrin and prunasin has afforded the opportunity of testing, by means of these cyanogenetic glucosides, for the presence of enzymes in plants capable of decomposing them, and liberating the HCN which they contain. Investigation of *E. maculata* showed that the sample of dried plant employed contained practically no enzyme, and that, therefore, only a trace of HCN was liberated, and, in spite of the presence of an exceptionally large amt. of glucoside, there still remained some doubt whether this plant was poisonous or not. Another plant, *Heterodendron elaeagnifolium*, recently submitted for examn., similarly contained a large amt. of cyanogenetic glucoside, but insufficient enzyme to give a positive reaction until enzyme from another source had been added. In this particular case it has been recorded by Petrie that only $\frac{1}{2}$ the amt. of acid was liberated before the addn. of enzyme (emulsin). Some of Petrie's work has been repeated, notably that involving amygdalin and its splitting enzyme β -amylgalase. The ease with which prunasin and sambunigrin are decomposed, as compared with amygdalin by the enzymes of plants, is shown, and the significance of the presence of enzymes in fodder plants is illustrated.

W. O. E.

Membranes of spores and pollen. II. *Lycopodium clavatum* L. 2. FRITZ ZETZSCHE AND HANS VICARI *Helv. Chim. Acta* 14, 58-62 (1931), cf. C. A. 22, 2949.—The sporemma from *lycopodium* spores previously described was not free from cellulose and ash. The cellulose could not be entirely removed by Schweitzer's cuprammonium reagent. Treatment with 85% H_3PO_4 gave a cellulose-free product, so d.d. 75% H_2SO_4 , but some sulfonation took place. Ash was reduced to < 0.2% by cooled HF. Hydroxyl groups were detd. by acetylation. The formula is fixed as $\text{C}_{20}\text{H}_{16}\text{O}_4(\text{OH})_4$. The yield was 23.8% of the spores. III. 2. *Picea orientalis*, *Pinus silvestris* L., *Corylus avellana* L. *Ibid.* 62-7.—Qual. tests on many pollens showed that their hulls or membranes are similar chemically to *lycopodium* sporemma, and the name *pollenar* is proposed for such compds. Three were isolated by treating the pollen successively with volatile solvents, KOH, H_3PO_4 , and HF, and their formulas detd. as: *Picea pollenar*, $\text{C}_{20}\text{H}_{16}\text{O}_4$; *Pinus pollenar*, $\text{C}_{20}\text{H}_{16}\text{O}_4$ or $\text{C}_{20}\text{H}_{16}\text{O}_4(\text{OH})_2$; *Corylus pollenar*, $\text{C}_{20}\text{H}_{16}\text{O}_4$. The yields were, resp., 21.9, 20.0 and 7.3% of the pollens. IV. 3. Fossil sporepollenar from Tasmanite and Russian lignite. FRITZ ZETZSCHE, HANS VICARI AND GUSTAV SCHÄFER. *Ibid.* 67-78.—Fossil spores and pollens have been found in many coals and oil shales. Selected specimens of a Tasmanian oil shale, tasmanite, were ground and the organic material was sep'd. from the sand by floating on CCl_4 . Clay was removed by shaking with soap soln. The crude sporepollenar, i. e., mixt. of spore and pollen membranes, obtained was purified as previously described. The yield was 17% of a substance with the probable formula, $\text{C}_{20}\text{H}_{16}\text{O}_4(\text{OH})_2$, for which the name *pollenar* is proposed. From a lignite from the Tschulchowsky deposits near Skopin

C_4 of a fossil spore in $C_{10}H_{16}O(OH)_2$ was isolated. This is believed to be derived from microspores of the genus *F. boissendianum* hence the name *boissendianin* is proposed for the compd. A Secotry spore was also present. Recent lycopodium spores and fossil bothrodendrin dissolve in concd HNO_3 in 12-24 hrs at room temp., whereas n requires weeks. This is attributed to the decreased no. of OH groups making it more paraffin like in character. L. E. GILSON

Biochemistry of tomato pigment. H. von ELZER, P. KARRER, E. von KRATTS and O. WALKER. *Helv. Chim. Acta* 14, 174-82 (1931).—Green tomatoes kept at 20-1° ripened normally in a few days, developing the usual red lycopodium coloring. At 30° they developed only a yellow color. Tests showed carotene, xanthophyll and an unidentified yellow flavone dye but no lycopodium. At 35° they turned yellow and spoiled rapidly. Light or its absence has no effect on the rate of coloring. It is believed that lycopodium is formed by an enzyme action which is inhibited by the higher temp. The pigment of the naturally yellow variety of tomato "Golden Queen" was studied. Dried skins from vine ripened fruit contained 0.021 g. carotene and 0.0033 g. xanthophyll per kg. but no lycopodium. A considerable amt. of a yellow flavone dye, insol in CS_2 sol in alk. was also present. This has not yet been identified. L. E. GILSON

The significance of iodine in plant and human metabolism. J. STOKLASA. *Z. Ernähr.* 1, 3-1 (1931).—I is absorbed by many plants in comparatively large quantities and changed to org. I. The I said plants have an important pharmacol. significance and can well be used in place of org. I. The I aids the oxidation processes in plants and animals and hinders a rise in H ion concn in the individual organs. The I absorbed by plants accumulates chiefly in the chlorophyll rich portion of the plant. For expts. on the absorption of I by plants and its distribution in org. and inorg. form, the following data are given. The pH of the soil was 6.5-7.0. High org. and inorg. soil I were closely correlated with high org. and inorg. plant I in the cases of *Cucumis sativus*, *Solanum lycopersicum*, *Euphorbia scabra*, *Phenanthus indicum* and *Spiraea alba*. For example, the fruit of *C. sativus* gave 120 parts per billion of total I with a soil content of 2.1 and 2.40 parts per billion with a soil content of 539. The pH values of the juices from the roots of various plants are given. By I fertilization the pH is raised. The author after 14 days on a diet of plants high in I, noticed a lowering of the H ion concn of the urine from pH 4.5-5.2 to 5.6-6.2. By the feeding of silk worms with leaves high in I a 61-55% increase in the no. of eggs laid was obtained. Thirty three references are given. FRANCIS P. GRIFFITHS

Organic matter given off by algae. ALBERT KROGH, EUGEN LANGER and WILLIAM SMITH. *Biochem. J.* 24, 16-20 (1930).—Culture expts. show that the org. material synthesized by the assimilation of freshwater algae is almost quantitatively stored in the cells of the algae. BENJAMIN HARKOW

Determination of maltose in plant extracts by maltase. NADIGAR NARASIMHA MURTI and MOTNARALI SREENIVASAYA. *Biochem. J.* 24, 1734-6 (1930).—The enzyme ext. was prepd. from finely ground brewer's yeast deaerated over P_2O_5 by treating the material with 10 parts of a 5% soln. of Na phosphate at 20° for 14 hrs. The ext. was centrifuged for 10 min. the filtrate neutralized to litmus with 4% KH_2PO_4 and preserved at 0°. This maltase ext. can be used for estg. maltose in mixts. of sugars occurring in plant exts. and tissue fluids. BENJAMIN HARKOW

Accumulation of electrolytes in plant cells. A suggested mechanism. G. F. BRIGGS. *Proc. Roy. Soc. (London)* B107, 248-50 (1930).—A possible scheme, based on alternating phases, one of greater permeability to anions and the other of greater permeability to cations is suggested to explain the accumulation of ions. J. S. H.

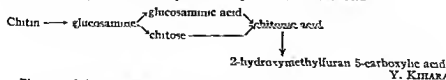
Alteration of the iron in chloroplasts due to chemical treatment. HANS GRIESSMAYER. *Planta Abh. E. Z. Wiss. Biol.* 11, 331-58 (1930).—When leaves of certain cereal plants were exposed in darkness to such vapors as SO_2 and fuming HNO_3 and to vols. of KCN and NH_4CNS the sol Fe in the plastids increased at the expense of the insol Fe. Plasmolytic tests with KNO_3 immediately after treatment showed that the leaf cells had not been killed. In contrast, treatment with ether and phenylurethan failed to increase the sol Fe. Placing leaves in an Fe soln. increased the Fe content of the plastids. Leaves lacking a normal green color due to etiolation, variegation or to lack of Fe all contained a normal amt. of Fe in the plastids. But after treatment with the chemicals previously mentioned, these abnormally colored leaves behaved the same as normal green leaves in regard to an increase in sol Fe. These results are offered as direct evidence for the participation of Fe in photosynthesis. A. E. H.

Experiments on the influence of iron salts on the pigments of chlorophyll. T. N.

GODNEV, S. K. KORZHENEVSKII AND M. H. GONCHARIK. *Ann Weissruth Staatl Akad. Landw Gory Gorki* 9, 126-35(1929)—In expts with *Pisum sativum* in culture solns it was found that there was no proportional relation between the Fe taken up from the soln and the coloring matter in the plants as was the case in maize. Fe is not of immediate importance for the building of chlorophyll, it is rather important primarily in the metabolic activity of the plasma. Fe enters as a constituent part of the chlorophyll mol. J. S. JOFFE

Limiting factors in photosynthesis. V. A. CHESNOKOV AND E. N. BAZURINA. *Comp rend acad sci U R S S* No 8, 193-8(1930A)—Ch and B analyze the theory of Blackman (*Ann Botany* 19, 28(1903)) and *Proc Roy Soc (London)* B76, 402(1905)) on the influence of direct and indirect factors on photosynthesis as it is based on the Liebig law of minima. They question the character of the optimum curve as expounded by Blackman. An increase in the temp of the medium which surrounds the leaf, according to Blackman, causes a series of changes in photosynthesis. Two groups of changes are noted, one corresponds to an increase in photochem activity following the van't Hoff reaction. The other one has to do with some processes in the plasma. Both influences take place simultaneously and the curve is a result of that. They point out that the temp of the leaf itself and not the external temp of the medium is responsible for the thermal increase in reactivity. In general they consider that all the factors in photosynthesis are nothing more than indirect factors. From the expts of Warburg and Negelein (cf *C A* 16, 3929-30) it is apparent that the coeff of utilization of light is very high even with a small light intensity, which shows that light influences other processes which in turn change the velocity of photochem reactions. As an increase in the temp of the leaf takes place, there is an increase in transpiration, dehydration of the plasma, a change in the penetration of CO_2 in the plasma, etc. The only rational way of investigating external factors is to det. the direct limiting factors. Such factors may be found in the internal system of the plant. They criticize the work of Lundegårdh on the effect of CO_2 concn and intensity of light on photosynthesis. They do not consider the CO_2 or light as the limiting factors. It is the velocity of the penetration of CO_2 in the plasma that is responsible for the changes observed. They explain Lundegårdh's expts as follows: Under the influence of light an increase in the penetration of CO_2 in the protoplasm takes place. This causes the chloroplasts to take up more moles of CO_2 and photosynthesis is speeded up. The same takes place with an increase in the concn of CO_2 . Thus the min is the CO_2 entrance in the plastids. They prove their point with the results described by Warburg. They conclude that Lundegårdh's new theory does not overthrow the fundamentals of Blackman, if the latter's theories are analyzed in the light of the authors' contention that it is the internal influences which are responsible for the behaviors noted. J. S. JOFFE

The fermentation products of molds. VIII. *Aspergillus glaucus*. YUSUKI SUMIKI. *J. Agr Chem Soc Japan* 6, 1153-65, *Bull. Agr Chem Soc Japan* 6, 105-6 (1930), cf *C A* 24, 3812—In the culture of *A. glaucus* 2 hydroxymethylfuran 5 carboxylic acid (m 163°) was produced and identified as the mono Ac deriv. (m 115°). It was not obtained in the culture solns of gluconic acid and hydroxymethylfurfural, but glucosamine, glucosaminic acid and chitose or chitonic acid were. The source of 2 hydroxymethylfuran 5 carboxylic acid may be chitin in the mold.



Phytin and the methods of its preparation. N. E. ZELIGSON. *Trans Sci Chem-Pharm Inst (Moscow)* No 23, 4-81 (in German 82-3)(1930)—A monograph contg a review of the literature, a restatement of the most reliable methods for the estn of the constituent element of phytin and a description of a new method for the double combustion of C. A new method is offered for the estn of the mineral phosphoric acid in phytin by pptn with molybdic soln in the cold. The following detns were made: total phosphoric acid, mineral phosphoric acid, phosphoric acid added to phytin previously freed from mineral phosphates, phosphoric acid in org combination, C and H, bases (Ca and Mg, Na and K, NH_4), Fe and Mn, Na and moisture. Analysis of several samples of phytin and a study of the Ba, Ca, Na, strychnine and borsx salts of inositol phosphoric acid were made. Inositolhexaphosphomolybdic acid was given

some consideration analytically. Summaries in Russian and in German are appended. B. S. LEVINA

Physico-chemical studies of rice produced in Gifu province. HANFUMON ITO, TATSUO WATANABE AND MASUO KITAHARA. *Gifu Imp Coll of Agr (Japan), Research Bull* 13, 48 pp (1930). —Rice with high protein content is generally harder. The protein content of rice of poor quality is generally high, but grain of superior quality is high in fat. The fat of rice varieties adapted to growth in colder regions is higher in acid value than the varieties requiring warmer climate. Elementary analyses of the gluten and globulin of rice showed the following values

	C %	H %	N %	S %	Ash %	Isoelectric point, pH
Glutelin	50.6-53.3	7.0-7.7	16.5-17.3	0.63-0.64	0.31-0.70	6.2-5.0
Globulin	50.8-52.5	7.0-7.6	16.5-17.3	0.44-1.15		5.5-4.6

K. KITSUTA

The role of iron and copper in the growth and metabolism of yeast. C. A. ELVEHJEM. *J Biol Chem* 90, 111-32 (1931). —The yeast was grown in a modified Wilder's medium the constituents of which were purified so that they were practically Fe- and Cu free. Addn of a small amt. of Fe accelerates the rate of yeast growth, increases the cytochrome content to the amt. normally present in com. yeast, and increases the respiratory quotient to Q_{O_2} values of 35-40. Addn of both Cu and Fe causes a further increase in the growth rate and the production of cytochrome with a distinctly higher α component. Cu has the property of stimulating the formation of certain hematin compds. 0.02 mg of Cu per 200 cc of medium is the optimum amt. for stimulating the formation of cytochrome α and additional amts. inhibit cell growth during the 1st 48 hrs. The toxic effect disappears as the no. of cells per unit of medium increases. The ratio of available Fe to the amt. of yeast present determines the Fe content of the yeast. Yeast low in inorg. Fe is unable to assimilate Fe from a phosphate buffer at pH 7.0 but readily takes up Fe at pH 4.0. Thus, the availability of Fe may be the limiting factor in many of the synthetic media which have been used for studying the growth requirements of yeast. Q_{O_2} values of 35-40 in 48-hr. cultures decrease to 15-20 when the culture is 6-7 days old because of the high acidity of the medium. At least 80% of the respiration is inhibited by KCN at 48 hrs. of age but the KCN labile portion gradually decreases and after 7 days the O_2 uptake is completely unaffected by KCN. The KCN stable system metabolizes glucose, is destroyed by heating to 60° and is inhibited by urethan to the extent of 70%. The respiration of yeast grown on beer wort, which prevents the development of high acidity, is largely inhibited by KCN. The total respiration of yeast grown on wort may be low on account of an insufficient supply of Fe in the beer wort. Bios stimulated the growth rate to a small extent but had no effect on the respiratory activity. The activity of bios is not so noticeable when the medium is maintained at the proper pH . Further work will be required to det. whether the effect of bios can be eliminated entirely by regulation of acidity or whether it actually stimulates yeast growth but in either case the effect will be closely associated with the availability of the Fe. Bibliography of 23 references.

A. P. LOTHROP

The non-volatile organic acids of green tobacco leaves. HUBERT B. VICKERY AND GEORGE W. PUCHER. *J Biol Chem* 90, 637-53 (1931). *Connecticut Agr Expt Sta Bull* 323, 155-202. —The existing methods for the detn. of the org. acids in plant tissues have been so modified as to make possible approx. quant. isolation and positive identification of the predominating acids. The acids are initially pptd. as Ba salts in the presence of dil. alc. rather than as the Pb salts by basic Pb acetate. In tobacco leaf exts. only about 1/3 of the total titratable acidity consists of acids of the familiar types such as malic, citric, succinic and fumaric. The remainder is due to acidic substances, the nature of which has not been detd. although part appears as phenolic acids. On account of the presence of these substances, indirect methods for the detn. of oxalic, citric and malic acids, such as those of Kissling and of Fleisher, do not yield trustworthy results. Bibliography of 29 references.

A. P. LOTHROP

Causes of unequal resistance of a vine to mildew. L. RIVES AND G. KALÉ. *Prog. agr.* 95, 138-9 (1931). —An examn. of the f. ps. of extd. juices of grape vine leaves, sampled in October, indicated no relation between their resistance to mildew and the concn. of their cellular juice.

F. W. MARSH

Variations in the osmotic pressure of certain Basidiomycetes. SYLVIA COLLA AND ZIFORA DANIN. *Boll. soc. ital. biol. sper.* 5, 885-7 (1930). —The juice of *B. edulis*,

A. cesarea L., *C. cibarius* Fries and *Russularrescens* Schaft was obtained by pressing and the Δ detd. The Δ varied from species to species it varied during the development of the reproductive cells and decreased during maturation and when aged, it was not equal in all parts of the same fungus
PETER MASUCCI

Respiration of seeds from oil plants. A I ERMAKOV AND NICOLAI N IVANOV *Biochem Z* 231, 79-91(1931) —The respiration of oil seeds during the early stages of germination does not differ from that of seeds contg carbohydrate stores. The CO_2 production bears a ratio of 1:1 to the O_2 absorption of these oil seeds. Neither the quantity nor the quality of the oil changes in the first hrs of germination, while the amt of sugar increases at the expense of polysaccharides undergoing enzymic hydrolysis. The fact that the I no. of the oil does not change leads to the conclusion that the oil is not directly oxidized by the O_2 from the air carbohydrates being the material primarily used in the respiration
S MORCULIS

The presence of a dihydroxyphenyl derivative and of a specific enzyme in the apple and other fruits. KARL SPITZER *Biochem Z* 231, 309-13(1931) —The presence of an *o*-dihydroxyphenyl deriv. has been established in apples, pears and other fruits, and its identity with dopa seems very probable. Also an enzyme has been found in these fruits which reacts in a sp way on dopa but has no effect on tyrosine. The cozyme is attached to the tissues. The browning of the fruit is due to its action on the chromogen
S MORCULIS

Relations existing between the different organic acids elaborated by *Sterigmatocystis nigra*. MOLLIARD *Compt rend* 192, 313-5(1931) —When *Sterigmatocystis nigra* is cultivated in media sufficiently low in minerals to insure the utilization of a given quantity of org substance, org acids are formed of which the most important are gluconic, citric and oxalic acids. From cultures grown for 90 days in a medium contg sucrose and only $1/20$ of the normal amts of mineral the following information was obtained. An appreciable amt. of reducing sugar remained at the end, at no time did oxalic acid appear, gluconic acid increased rapidly for 10 days, remained stationary from the 10th to the 60th day and then rapidly diminished, citric acid was not very abundant at the end of 10 days, but increased in concn constantly throughout the entire 90 days. To det whether citric acid is formed from gluconic acid a culture of *S. nigra* was grown under conditions similar to those of the above expt, gluconic acid being used instead of sucrose as the base. The following results were obtained: No trace of citric or oxalic acids was formed; the polarimetric deviation remained proportional to the free acidity. Gluconic acid, then, acts as a nutrient in the same manner as glucose, but never gives rise to citric acid. Gluconic acid, like the sugars, gives rise to oxalic acid when the culture medium tends toward alkali. Analogous conclusions are reached concerning citric acid. M. concludes that gluconic and citric acids must be considered as oxidation products of sugars because of inactivity in mineral substances, they both may be utilized subsequently by the fungus, but citric acid cannot be formed from gluconic acid. Oxalic acid is incapable of being utilized by *Sterigmatocystis nigra* and must be considered a final excretion product.
ALBERT L. RAWLINS

The influence of nickel compounds and of cyanides on the germination of grains. ANNELIESE NIETHAMMER *Wiss Arch Landw, Abt A, Pflanze* 4, 607-34(1930) —Grains of wheat were placed in vessels contg H_2O , NiSO_4 , $\text{Ni}(\text{NO}_3)_2$, NiCl_2 , NiCO_3 , $\text{Ni}(\text{CN})_2$, 2KCN , KCNS , NaCNS solns. of 0.1, 0.5 and 1.0% and the germination was observed. Arranged in the order of injurious effects, the compounds rank $\text{Ni}(\text{CN})_2$, $2\text{KCN} > \text{Ni}(\text{NO}_3)_2 > \text{NiCO}_3 > \text{NiCl}_2 > \text{NiSO}_4$, 0.1% solns. of $\text{Ni}(\text{NO}_3)_2$ and NiSO_4 accelerate germination. It was not detd whether the effect was due to physical or chemical changes, although histological examn showed that the Ni solns penetrated to the embryo as well as the endosperm. Wheat grains of the years 1921, 1923 and 1924 were similarly treated, and similar results were observed. The individual grains were then examd quantitatively for acetaldehyde, starch, reducing sugar and protein. The same investigation was conducted on *Allium cepa*
W GORDON ROSE

The algae flora of the Ruhr River. HERMANN BUDDE *Arch Hydrobiol* 21, 559-648(1930) —Numerous data on the compo. of the Ruhr water are included. The analyses comprised temp., transparency, color, sediment, reaction to phenolphthalein, H ion concn, hardness, no. of germs and content of org substance, Cl, HNO_3 , HNO_2 , NH_4 , Fe, CO_2 and O_2 . Large bibliography.
G. SCHWACH

The metabolism of mosaic-diseased and healthy tobacco plants. M LÖDTER *Phytopath. Z* 2, 341-59(1930); *Rev. Applied Mycol* 10, 132-3 —Chem. analysis of healthy and mosaic tobacco plants showed that the accumulation of starch in the latter is not due to a disturbance or reduction of diastatic activity. A sucrose- and a pep-

respiratory quotient for the 38th to 44th hr. inclusive, is 0.733, uncorrected, or 0.722 if corrected for protein metabolism. The gas-exchange studies were definitely controlled and can be reproduced. C. R. FELLERS

Specific dynamic action. GRAHAM L. K. J. *Nutrition* 3, 519-30 (1931) — A review. L. believes that sp. dynamic action of carbohydrate is due to heat production of the intermediary metabolism rather than to carbohydrate plethora. His views are now in agreement with those of Rubner. C. R. FELLERS

Investigations in the feeding of cottonseed meal to cattle. J. O. HALVERSON AND F. W. SHERWOOD. North Carolina Agr. Expt. Sta., *Tech. Bull.* 39, 5-158 (1930) — Eight years' results are reported in full. The feeding of large amounts of cottonseed meal and hulls to cattle over considerable periods of time results in what is commonly called 'cottonseed meal poisoning', a condition formerly ascribed to a toxic substance, *gossypol*. Numerous expts. showed definitely that the disease was not due to a poisonous principle but rather to nutritive deficiencies of the rations. A crit. study of the various combinations of nutritive adjuncts used indicates that cottonseed meal and cottonseed hulls do not contain sufficient Ca or vitamins A, D or B to meet fully the needs of the animals. Great care must be used to select rations furnishing adequate amounts of minerals and vitamins for dairy cows. White corn is likely to be deficient in Ca, P and vitamin A and corn silage or grain should not be used to supplement cottonseed feeds. The typical symptoms of cottonseed poisoning, *s. e.* staggers, impaired vision and spasms, were readily reproduced on heifers receiving meals other than cottonseed such as linseed, peanut and soy bean in the presence of minerals. On autopsy, the livers of the animals were shown to be almost totally lacking in vitamin A. Cod liver oil added to the ration proved moderately efficacious in preventing or curing the condition. Reproduction of cows fed on a cottonseed or similar deficient ration was decidedly subnormal. Numerous expts. where cottonseed meal was supplemented by poor roughage, wheat straw and a small quantity of green feed showed inadequate maintenance, poor reproduction and other conditions attributable to Ca and vitamin A deficiency. Similarly, rations of cottonseed meal contg. minerals, timothy hay and corn silage did not quite maintain health and vigor in the animals. Even when $\frac{1}{2}$ of the entire ration was alfalfa hay and beet pulp was fed in addn., the ration was inadequate for dairy cows over an extended period of time. Cottonseed or other meals from high oil bearing seeds must be freely supplemented with good hay and possibly non maize grains to avoid various nutritive deficiencies. Bibliography. C. R. FELLERS

Corn silage versus a mixture of wet beet pulp and molasses for milk production. C. F. MONROE, C. C. HAYDEN AND A. F. FRANKS. Ohio Agr. Expt. Sta., *Bimonthly Bull.* 148, 3-8 (Jan.-Feb. 1931) — Slightly higher milk and butterfat production were obtained on the beet pulp-molasses ration, though insufficient to justify the increased cost of the ration. C. R. FELLERS

Does the feeding of corn silage to dairy cows lead to acidosis? A. E. FRANKS, C. C. HAYDEN AND C. F. MONROE. Ohio Agr. Expt. Sta., *Bimonthly Bull.* 148, 8-11 (Jan.-Feb. 1931) — The total acidity of corn silage averaged 2-3%, consisting of lactic and acetic acids. The blood and urine of cattle normally have a slightly alk. reaction. The urine of several cows fed daily rations contg. 30-50 lb. of corn silage as compared with beet pulp showed no significant increase in either total acidity or H⁺ ion concn. NH₃ and bicarbonate CO₃ were also unaffected by the feeding of silage. Conclusion: The org. acids in silage are fully metabolized and do not appear as org. acids or related compds. in the urine. C. R. FELLERS

Relation of food to the growth of pre-school children. MARY A. BROWN. Ohio Agr. Expt. Sta., *Bimonthly Bull.* 149, 73-86 (Mar.-Apr., 1931) — Two groups were formed based on different economic levels, of 114 kindergarten children. The children were regularly weighed and examd. for phys. defects, etc., for approx. 1 year. Conclusion: It is desirable to include a quart of milk, vegetables, cooked and uncooked fruits, and cereals in the daily diet of the growing child, and eggs should be used several times a week, since a greater rate of growth was observed for those children whose diets contained these foods than for those children whose diets were lacking in quantity in certain food elements. C. R. FELLERS

Adsorption of vitamin A on silica gel. L. L. LACHAT, R. ADAMS, DUTCHER AND H. E. HONEYWELL. Penna. Agr. Expt. Sta., *Bull.* 258, 8 (1930) Abstract; 433 *Ann. Report of Director* — Vitamin A in cod liver oil may be adsorbed on highly activated silica gel so tenaciously that it cannot be rendered available when the silica gel is fed to rats. Toluene removes the adsorbed vitamin A from the silica gel, while acetone extracts are inactive. Further studies on availability before and after adsorption are being continued. C. R. FELLERS

Digestibility by sheep of the constituents of the nitrogen-free extract of feeds. G. S. I. RAPS. Texas Agr. Expt. Sta., *Bull.* 418, 5-15(1930).—The contents of starch, pentosans, residual N free ext. and sugars were detd. on 56 feeds. The digestibility by sheep of the sugars, starches, pentosans and residual N free ext. was detd. by animal feeding tests. Feeds known to be of high feeding value are characterized by a high content of starch, or starch and sugar combined, or if the starch content is not high by a high content of protein. The N-free ext. of feeds of low feeding value contains a high percentage of pentosans and of residual N free ext. The N free ext. of some feeds contains an appreciable percentage of compds. which are not carbohydrates, whose digestibility is uniformly high. For several feeds the digestibility of pentosans was 50-60%, but was somewhat variable. The pentosans in crude fiber are usually digested to a less extent than the total pentosans, or the pentosans in the N free ext. The residual N free ext. in hays and fodders is generally digested to a greater extent than are the pentosans. C. R. FELLERS

The nature of dietary deficiencies of milk. J. TRAVESTINE BECKER AND E. V. MCCOLLUM. *Am. J. Hyg.* 12, 503-10(1930).—Of 38 diets in which milk powder was used at 60-85% levels in combination with other substances, only 24 sufficed for the production of young rats. The best supplements from the standpoint of reproduction records were cooked dried beef liver, yeast and a mixt. of ferric citrate and copper sulfate. The liver is believed to owe its value to the Cu and Fe which it contains. S. B. FOSTER

Effect of cod-liver oil on calcium metabolism of young chicks. ARTHUR D. HOLMES AND MADLEINE G. PIOTT. *Ind. Eng. Chem.* 23, 190-6(1931).—Four hundred vigorous, day-old Rhode Island Red chicks were divided into 12 comparable pens, and the effect of cod liver oil on their Ca utilization was detd. (1) when the Ca was obtained from different sources, (2) when it was fed at different levels, and (3) when it was fed at varying ratios to P. Judged by the rate of growth and bone development, the chicks which received cod-liver oil showed better Ca utilization than the controls. The oil was most effective when the Ca-P ratio was low. When this ratio was 5-1 the growth and Ca utilization of both exptl. and control groups were unsatisfactory. S. B. FOSTER

Radiographic study of the effect of irradiated ergosterol on the healing of experimental fractures. J. MORELLE. *Rec. Belge Sci.* 2, 226-8(1930); *Bull. Hyg.* 6, 83(1931).—M. studied the action of vitamin D on the lesions of the fibula of rats. The rats were divided into 3 groups which were given 0.001 mg., 0.01 mg., and 2-4 mg. of ergosterol per day. Up to the 20th day no results were visible, but after that time in those receiving the medium dose the fracture lines were invisible, while in the controls they were still apparent. After the 30th day there was a dense callus in treated animals while in the controls the disappearance had just begun. Heavy doses caused abnormally dense and exuberant callus formation. GEORGE R. GREENBANK

The mode of action of the vitamin B complex. I. EMIL ABERHALDEN. *Arch. ges. Physiol.* (Pflügers) 226, 723-37(1931).—The paralytic symptoms of pigeons fed on a diet of polished rice could be abolished by feeding small quantities of various organ exts. (particularly liver). The bearing of these expts. on the problem of the nature of the vitamin B complex is discussed. II. EMIL ABERHALDEN AND V. VLASSOPOULOS. *Ibid.* 808-15.—The dehydrase content of liver, muscle, and brain tissue of pigeons maintained on a diet of polished rice is abnormally low. The addn. of yeast or of dried organ exts. in the diet of such pigeons restores the above mentioned organs to their normal dehydrase content. ARTHUR GROLLMAN

The utilization of soy-bean flour and a new soy protein preparation by man and the lower animals. J. KAFFHAMMER AND H. HABE. *Deut. med. Wochschr.* 56, 1168-70(1930).—Metabolism expts. on men and dogs showed the value of soy-bean flour as a source of protein in the diet. *Novo-Tropen*, a com. prepo. contg. soy bean flour, also proved efficacious as a source of protein. ARTHUR GROLLMAN

Carotene and vitamin A. L. K. WOLFF, J. OVERHOFF AND M. VAN ECKELEN. *Deut. med. Wochschr.* 56, 1428-9(1930).—Carotene and vitamin A were sep'd from a no. of plant and animal substances. Their detn. in such substances is described. ARTHUR GROLLMAN

The effect of hypophyseal extracts and various protein bodies on the specific dynamic action of protein. ERNST HERZFELD. *Deut. med. Wochschr.* 56, 1558-60(1930).—Muscle, milk, and liver protein showed a specific dynamic action in normal individuals. ARTHUR GROLLMAN

"Diet 4" for breeding rats for work on vitamin A. SKULI V. GUDJONSSON. *Biochem. J.* 24, 1591-4(1930).—The diet contains (in per cent) skimmed-milk powder

Simple methods for metabolic balance studies and their interpretation. EVA DONELSON, BETTY NIMS, HELPN A. HUNSCHER, CARROLL F. SHUKERS AND ICIE G. MACY. *J. Home Econ.* 23, 267-72(1931).—The mean difference between the chem. compn. of a diet as computed from standard tables and as detd. in the lab was 2-5% N varied 0.7-38%, Ca varied 0-22%, and P varied 0-38%. The magnitude of retention or loss that can be considered significant in any expt is detd. by calcg the av. deviation. A calcn. of metabolic balance for short observation periods shows daily fluctuations, while balances for longer periods show the general tendency better. A list of 23 references is given. AMY LEVESCONTE

The effect of feeding sorghum and sugar-cane sirups on nutritional anemia. OLIVE SHEETS AND ERNESTINE FRAZIER. *J. Home Econ.* 23, 273-80(1931).—Sorghum and sugar-cane sirup are effective in producing hemoglobin in the rat. A few expts with Crooked Top sirup proved it to be almost as potent as calf liver. A. LEV.

A differentiation of the so-called antipellagic factor, vitamin G. BARNETT SURE, MARGARET ELIZABETH SMITH AND M. C. KIK. *Science* 73, 242-3(1931).—Many expts indicate that vitamin G contains 2 factors, one producing dermatitis without affecting the growth of rats, the other stunting the growth without sign of skin lesions. It is suggested that the growth promoting factor be designated by the letter F. AMY LEVESCONTE

Comparison between irradiation of diet and supplemental irradiation of animals in vitamin A and D deficiency. F. E. CHIDESTER, A. G. EATON AND N. K. SPEICHER. *Science* 73, 190-1(1931).—Rats depleted of vitamins A and D were fed on a diet (Sherman No. 380) low in D and deficient in A and given 0.01 mg. of ergosterol daily. Some of these rats were irradiated also for 30 sec. daily. The addnl. irradiation stimulated growth for a short time but did not produce the striking effects observed when Fe₂ was also added. I. M. LEVINE

Sterol content and antirachitic activity of mold mycelia. L. M. PREUSS, W. H. PETERSON, H. STEENDOCK AND E. B. FRED. *J. Biol. Chem.* 90, 309-84(1931).—The autoclaved, dried and finely ground pads, resulting from the growth of certain molds in a synthetic, inorg. medium contg. 4% glucose as the source of C, were irradiated with ultra-violet light and fed to rachitic rats. Five kinds of mushroom were also tested for antirachitic activity. Eight of 12 molds and 3 of the mushrooms brought about distinct healing of rickets even when fed at a level of 10 mg. of irradiated material per rat for 7 days. The total alc. sol. sterol, detd. by the digitonide method, in 11 molds and 4 mushrooms varied from about 0.1 to 1.0% of the dry wt. of the fungus material. In most cases 90% or more of the alc. sol. sterol was present in the free state. The sterol content of the molds varied not only with the species but also with different strains of the same species and was decreased by increasing the glucose content of the medium. Supplementing the medium with such org. nutrients as an aq. ext. of malt sprouts or fresh yeast did not appreciably change the sterol content but the mycelial growth of 1 of the strains studied, *Aspergillus oryzae*, was increased from 20 to 50%. A decrease of 50% in the wt. of the pad resulted from increasing the incubation period of this strain from 10 to 50 days. In most cases increasing the glucose content of the medium to 10 or 20% more than doubled the wt. of the dry pad. Bibliography of 22 references. A. P. LOTHROP

The induction of tetany in rachitic rats by means of a normal diet. ALFRED F. HESS, MILNRED WEINSTOCK, H. R. BENJAMIN AND J. GROSS. *J. Biol. Chem.* 90, 737-46(1931).—"Tetany can be induced in rachitic rats simply by an abrupt change from a rickets-producing ration, high in Ca and low in P, to a normal ration of dried milk or of dried milk and whole wheat. The fall in Ca in the serum which is brought about by this means develops within 48 hrs. but is maintained for only a few days. This striking reaction is due not to an absolute or to a relative increase of P in the dietary but to a sudden shift in the Ca/P ratio in the subsequent diet as compared to the preliminary diet. When this ratio of Ca/P is decreased from about 4.1 to 1 or 1.5:1 tetany ensues, whereas if it is decreased only to about 2:1 the fall in Ca and nervous symptoms do not come about. Attention should be directed to the effect of marked alterations in the constitution of dietaries, as such shifts may help to explain nutritional disturbances which are inexplicable merely from the standpoint of adequacy."

The effect of heat upon the biological value of cereal proteins and casein. AGNES F. MORGAN WITH THE COÖPERATION OF FLORENCE B. KING, RUTH E. BOYDEN AND V. ANN PETRO. *J. Biol. Chem.* 90, 771-92(1931).—The digestibility of proteins toasted at 150-200° for 30-45 min. is but little different from that of the raw protein but there are differences in the biol. values (ranging from 10 to 19%) which are in all cases more

than 6 times as great as their probable errors. The unexplainable loss of N occurs chiefly in the urine, indicating that the change produced by the heat treatment lies probably in the assortment or availability of the amino acids absorbed. The addn. of 5% casein to the toasted diets very nearly made up the discrepancy between the latter and correspondingly supplemented raw diets, thus indicating that the deficiency observed lay in the protein fraction of the toasted diet. With wheat gluten the max. growth rate (1.4 g per g of protein eaten in 56 days) occurred when the diet contained 18% of the protein. At an approx. 12% level the biol. values of raw and toasted gluten were 60 and 54, resp. The values for raw and toasted casein at an 8% level were 66 and 53. Cooking with H_2O , in general, slightly increased the biol. value of the protein. The values of raw, water-cooked and toasted whole-wheat protein alone and supplemented with 5% casein were found to be 64, 67 and 52 for wheat alone and 67, 75 and 69 with wheat plus casein. Bibliography of 33 references. A. P. L.

Digestive erythropenia and factors which cause it. GIUSEPPE SOLARINO AND VITTORIO CATTAINO. *Boll. soc. ital. biol. sper.* 5, 985-9 (1930).—The administration of a diet rich in protein substances (meat) or in $NaCl$ in concn. corresponding with that of gastric juice causes in normal dogs and in man a marked, const. erythropenia. Erythropenia more marked and of longer duration follows the administration of liver. Digestive erythropenia is related to 2 series of factors: to the accumulation of blood in the splanchnic area and to the erythrocytic activity of certain hemopoietic tissues.

PETER MASUCCI

Studies on the metabolism of fats. VII. The glycogen content of adipose tissue of rats under varying conditions of alimentation. G. SCOR. *Boll. soc. ital. biol. sper.* 5, 1037 (1930), cf. C. A. 24, 5804.—Glycogen deins expressed as mg per 100 g. adipose tissue were: fasting animal 181, animal at const. wt. 527, animal after refeeding 1600.

PETER MASUCCI

The diet in diabetes of children. PEDRO B. LANDAUER AND FÉLIX PUCHTLU. *Semana med. (Buenos Aires)* 1931, 1, 360-3.—The caloric intake should be in the first decade of life 69 cal per kg. wt., in the second 48 cal. The requirement for carbohydrates is 2.09 g per kg. in the first decade and 1.44 g. in the second decade, for proteins it is 2.47 and 1.58 g., resp. and for fats 5.57 and 3.87 g. Only in a few cases is it possible to dispense with insulin.

A. E. MEYER

Calcium metabolism. ALBERT RANSOY. *Semana med. (Buenos Aires)* 1931, 1, 465-7.—After a short review on the subject, opocalcium is recommended, a product composed of minerals, parathyroid, suprarenal and thymus gland, which has been exposed to ultra violet light.

A. E. MEYER

Results of exclusively dietary treatment in forty-six cases of obesity. EUGENY MÖLLER. *Acta med. scand.* 74, 341-52 (1931).—Forty-six patients were treated with a diet contg. 930 calories with about 6 g. protein N per day. This treatment continued for 54 days, the av. loss in wt. sustained being 8.8 kg. The basal metabolic rate of one patient who lost 29 kg. in 5 months showed no significant changes during that period.

S. MOROGLIS

Influence of some fats and oils, meat and bread on growing rats. V. A. LEIROVICH-LYSSINA. *Biochem. Z.* 231, 260-73 (1931).—Large amts. of cod liver oil cause toxic effects which manifest themselves in extensive losses of body wt., trophic disturbances and premature death. With corresponding quantities of olive oil the body losses are smaller and the animals do not die prematurely even in long-continued expts. Small or moderate amts. of olive, linolic and other oils do not stimulate the growth of the rat. The greatest stimulating effect upon increase in body wt., especially following a preliminary fast, has been obtained with 5 g. meat added to a bread diet or with meat alone. The stimulating effect of the meat lasts longer in growing rats which have had a previous fasting experience than in rats without this.

S. MOROGLIS

The presence of vitasterol A in the cacao and its products of extraction. H. LAUBE, HEIM DE BALSAC AND R. LERAT. *Bull. soc. therap.* No. 10, 12 pp (1930), cf. C. A. 25, 729.—The following exts. of the cacao bean were made: (1) cacao butter, (2) alc. ext. of bean, (3) alc. ext. of shell, (4) the tailings of crystn. after extn. of α - and β -sterols. These products, in comparison with cod-liver oil, were tested by the color reaction with $AsCl_3$ and $SbCl_3$ according to the procedure of Carr and Price. The alc. ext. of the bean gives a feeble color reaction with both reagents, $AsCl_3$ and $SbCl_3$, the alc. ext. of the shell and the tailings of crystn. give color reactions comparable to that obtained with cod liver oil. These exts. were further tested for their growth promoting and antirachitic properties in expts. with white rats which had been sensitized by a vitamin A-deficient diet according to a procedure of Lecoq. It was shown that important amts. of a vitasterol, similar to vitamin A of growth and develop-

the acid fed animals increased in wt.; in the other groups the base fed animals gained. The av. gain of 27 base-fed animals was 8% greater than the av. of 27 acid fed animals. Diuresis caused in both acid- and base fed animals the same loss of body wt. which shows that the difference in growth was not caused by the retention of fixed fluid.

H. TAUBER

The significance of the calcium and phosphorus content of the food. K. WALTNER. *Magyar Orvos Archivum* 31, 276-82(1930).—For white rats the min. quantity of P necessary for optimum growth is the same as that of Ca. A food contg. P and Ca in the same ratio as that in human milk is deficient in P for rats. Twice the min. amt. of one of the elements can be ingested without injury in the presence of the min. amt. of the other. Ca deficiency of the food can be partly compensated by vitamin D, but the deficiency in P only to a very limited degree.

H. TAUBER

The influence of liver diet on the action of hemolytic poisons. V. FENK. *Magyar Orvos Archivum* 31, 446-50(1930).—The development of anemia in rats caused by phenylhydrazine poisoning is retarded by liver diet, but the regeneration of blood is not stimulated. Typical anemia cannot be produced by Cu acetate. Liver diet has no effect upon Cu acetate poisoning.

H. TAUBER

The destructive action of finely divided solids on vitamin A. JOSEPH K. MARCUS. *J. Biol. Chem.* 90, 507-13(1931).—The activity of vitamin A concentrates from cod-liver oil (Marcus, C. A. 23, 1162, 4093) is largely destroyed on standing 8-28 days in presence of various finely divided powders, whether in air or CO₂.

K. V. T.

The fundamental food requirements for the growth of the rat. VI. The influence of the food consumption and the efficiency quotient of the animal. LEROY S. PALMER AND CORNELIA KENNEDY. *J. Biol. Chem.* 90, 545-64(1931), cf. C. A. 23, 4499.—The digestibility of various stock diets by rats varies from 80 to 95%, and diets should therefore be compared on the basis of an equal intake of digestible dry matter. This is done very reproducibly by collecting, drying and weighing the feces. When diets are thus compared, differences in the growth curves on synthetic vitamin contg. and natural complete diets almost disappear, though the natural diet appears still slightly superior. Furthermore, the efficiency of food utilization, expressed as the efficiency quotient (E. Q.), or gain in wt. per 100 g. body wt. per g. of food digested, is higher in males than in females. The greater gain in wt. of males is due to this (cf. Osborne and Mendel C. A. 20, 3487; Mendel and Cannon, C. A. 22, 798). Differences in E. Q. also account for the variations in growth between groups of rats on the same diet. The stimulating effects of fresh lettuce, liver and carrots, given with yeast, on the growth of rats are shown to be due solely to effects on the food consumption, when allowance is made for differences in the E. Q. With the same method yeast and wheat embryo are found equally effective as sources of water sol. vitamin in a diet otherwise complete. By means of the E. Q., growth expts. can be more accurately controlled, and individual variation in gain in wt. of animals can be partially eliminated.

K. V. T.

A critique of the line test for vitamin D. CHARLES E. BILLS, EDNA M. HONEYWELL, ALICE M. WICKICK AND MILDRED NUSSMEIER. *J. Biol. Chem.* 90, 619-36(1931).—The 4 grades or degrees of healing of rickets devised by Bills and McDonald (C. A. 20, 2522) are not proportional to the dosage of vitamin D. A graphic method is described whereby the amt. of dilt. required for any vitamin-D-contg. prepn. to give 2+ healing can be detd. Av. cod liver oil is taken as that which induces 2+ healing when given to rickety rats at 1/4% in diet 3143 for 5 days. The sex, color and wt. of rats are not of importance in the line test. The relation of dosage required to the duration of the test period (between 3 and 14 days) is plotted, and the probable error of the method estd.

K. V. THIMANN

KLINKE, KARL. Der Mineralstoffwechsel. Physiologie und Pathologie. Band III of "Einzeldarstellungen aus d. Gesamtgebiet d. Biochemie." Leipzig: Deuticke. 298 pp. M. 24, linen, M. 27.

SHERMAN, H. S. AND SMITH, S. L. The Vitamins. 2nd ed., revised and enlarged. Am. Chem. Soc. Monograph No. 6. New York: The Chemical Catalog Co., Inc. 538 pp. \$6.

F—PHYSIOLOGY

R. K. MARSHALL, JR.

Basal metabolic rate of medical students and nurses in training at Charleston, S. C. ROSE E. REMINGTON AND F. BARTOW CULP. *Arch. Internal Med.* 47, 366-75(1931).—The basal metabolic rate of student nurses and male medical students at Charleston, S. C., showed results 10% lower than the Aub-Dubois standards. The I intake, dietary

habits and the state of nutrition are not believed to be responsible for the results obtained
J. B. BROWN

Effect of acidosis upon capacity for work. H. DENNIG, J. H. TALBOTT, H. T. EDWARDS AND D. B. DILL. *J. Clin. Investigation* 9, 601-13 (1931); cf. *C. A.* 24, 2504 — Acidosis produced by the oral ingestion of NH_4Cl lowers the capacity for work as evidenced by decrease in capacity for O debt. The buffer value of blood proteins is reduced $\frac{1}{2}$ or more with a decrease of $\frac{1}{2}$ in CO_2 capacity. The ability to neutralize lactic acid is lowered. These factors are raised in alkalosis.
J. B. BROWN

Normal standards of gastric function. W. SCOTT POLLARD AND ARTHUR L. BLOOMFIELD. *J. Clin. Investigation* 9, 851-8 (1931) — The results of gastric analysis, detd. under fasting basal conditions and resulting from the injection of histamine, are plotted in a number of charts showing the normal ranges. The charts are useful in clinical diagnosis.
J. B. BROWN

Studies on the physiology of the parathyroid glands. II. The relation of the serum calcium to the serum phosphorus at different levels of parathyroid activity. FULLER ALBRIGHT, WALTER HAVER, JESSIE REED COCKRILL AND REAR ELLSWORTH. *J. Clin. Investigation* 9, 659-77 (1931); cf. *C. A.* 23, 5231 — Study of a large no. of simultaneous detns. of serum Ca and inorg. P in patients with hypo- and hyperparathyroidism shows a slight reduction of P at high Ca level and a marked increase of P at low Ca level, so that the product of Ca and P is roughly const. Parathyroid dysfunction is a disorder in which body fluids show normal Ca phosphate but an abnormal ratio of Ca to phosphate. At high Ca levels, however, P is no longer excreted and the fluids contain high Ca phosphate, with calcification elsewhere than in the bones as the result. The data do not explain the casual relationship between Ca rise and P fall.
J. B. BROWN

The effect of injections of female sex hormone (estrin) on conception and pregnancy in the guinea pig. G. LOMBARD KELLY. *Surg., Gynecol. Obstet.* 52, 713-22 (1931); cf. *C. A.* 24, 3816 — Small doses of estrin injected into female guinea pigs prevent conception. Pregnancy may be interrupted by this hormone. An excess of female sex hormone in the blood over the corpus luteum hormone is incompatible with conception.
J. B. BROWN

Hormonal sterilization. K. JUNKMANN. *Z. ges. exp. Chm.* 44, 162-3 (1931) — A review and bibliography are given of the different chemotherapeutical ways of producing facultative sexual sterility in the animal and human body. During pregnancy the corpus luteum prevents the ripening of new ova by internal secretion. The same effect is obtained by the injection of preps. from ovaries or placentas of pregnant animals, such as phletoaptone (Merek), progynon, folliculin, menformone or the inner secretions of the corpus luteum. The hormones of the anterior lobe of the pituitary body can have the same effect, though the mechanism of their action is different. In the male body sterility can be produced by administering female sexual hormones.
ALFRED BURGER

Combined peptic-tryptic digestion in vitro and the digestibility of foodstuffs. A. RONCATO. *Arch. fisiol.* 28, 69-98 (1930) — See *C. A.* 25, 329.
E. J. C.

The existence of a pancreas hormone which lowers blood pressure. P. GREY AND N. KISTHINOS. *Wiener klin. Wochschr.* 43, 1530-6 (1930) — The amt. of the hormone is quite variable in different com. insulin preps. Its effect is not the same as that of histamine or choline. It neutralizes the effect of adrenaline on blood pressure.
D. B. DILL

Action of the heart hormone, "Eutonon," from the liver. H. SALOMON AND G. ZUELLER. *Z. ges. exp. Med.* 66, 291-324 (1929) — The response to eutonon was detd. in the frog heart, various heart lung preps. and in selected clinical cases.
F. L. DUNN

The antimasculine action of the female sex hormone, menformone. E. BORCHARDT, F. DINGEMANN, S. E. DEJONCHER AND E. LAFITTE. *Z. ges. exp. Med.* 68, 86-103 (1929)
F. L. DUNN

Corpus luteum and the sexual hormone. L. FRAENKEL AND E. FELS. *Z. ges. exp. Med.* 68, 172-81 (1929).
F. L. DUNN

A hormone affecting heart activity. XIII. Effect on the frog heart. L. HABERLANDT. *Z. ges. exp. Med.* 68, 185-95 (1929); cf. *C. A.* 24, 5817-8, *Med. Klin.* 1929, No. 14 — The action of Hormocardiol (Hoechst) on the frog heart is discussed.
F. L. DUNN

Effect of muscle activity, adrenaline and stimulation with the galvanic and faradic currents upon the blood sugar. M. DORLE. *Z. ges. exp. Med.* 69, 242-54 (1930) — Below the point of fatigue muscle activity produced a variable response in the blood sugar and in various diseases and constitutional types. Adrenaline produced the same reaction in the two constitutional types. Elec. stimulation of the gastro-intestinal canal

in animals produced the greatest rise in blood sugar. The rise in blood sugar did not parallel the rise in blood pressure following the injection of adrenaline. F. L. DUNN

The mercapto group in blood cells. P. GANNE. *Z. ges. exper. Med.* 69, 392-421 (1930).—A method is described based on the oxidation of the SH group by $K_4Fe(CN)_6$ in strongly acid soln. by which the relation of SH and SS groups in glutathione and thionine can be detd. in blood. In venous blood the SH form is found, while in arterial or oxygenated blood the proportion of SH group is reduced 10-50%. The SH combination remains attached to protein when the soln. is treated with CCl_3CO_2H , neutralized with U acetate in the cold, but it is sepd. following hemolysis. The SH content of the blood cells increases following bleeding, anoxemia and PhNH₂ poisoning. There is an extensive bibliography. F. L. DUNN

Metabolism at reduced pressure. I. H. ELIAS AND M. TAUBENHAUS. *Z. ges. exper. Med.* 69, 529-60 (1930). F. L. DUNN

Behavior of 8-keto- α -hexoic acid in the perfused liver. WALTER L. DULIERE AND HENRY S. RAPER. *Biochem. J.* 24, 1672-7 (1930).—The Na salt of 8-ketohexoic acid on perfusion through the liver yields aretome bodies, this indicates a 8-oxidation. BENJAMIN HARKOW

Male hormone. IV. CASIMIR FUNK AND BENJAMIN HARKOW. *Biochem. J.* 24, 1678-80 (1930). cf. *C. A.* 24, 5321.—A simplified method of prep. the hormone is described. This involves the extrn. of strongly acidified male urine with $CHCl_3$, heating the evapd. $CHCl_3$ residue with alkali, repeatedly extg. the alk. solution with ether, evapg. the ether ext. and taking up the residue with oil. About twice as much of the hormone can be obtained from a strongly acidified urine as from an alk. urine. An important difference in the chem. behavior of the male and female hormones has been noted. BENJAMIN HARKOW

The presence of a volatile organic chlorine compound in blood. NOAH MORRIS AND SAMUEL MORRIS. *Biochem. J.* 24, 1716-22 (1930).—For each blood there is a certain max. and min. amt. of volatile Cl present beyond which the content does not pass. The Cl in normal blood probably exists in 3 forms: (a) inorg., (b) org. non-volatile and (c) org. volatile. BENJAMIN HARKOW

Studies on the biochemistry of sulfur. VIII. Rate of absorption of cystine from the gastro-intestinal tract of the white rat. M. X. SULLIVAN AND W. C. HESS. *U. S. Pub. Health Repts., Suppl.* No. 89, 16 pp (1931), cf. *C. A.* 25, 2449.—Studies are made on the rate of absorption of cystine, fed as the Na salt, from the gastrointestinal tract of the rat by means of 3 cystine methods, the Folin Marenzeller method, the Okuda iodometric method and the Sulzbao method. The last two methods are in close agreement when corrections are made for the cystine value of the Na cystinate fed and the % recovery of cystine in the method of extrn. The value found by these methods is, in round nos., 60 mg. of cystine per 100 g. of body wt. per hr.—a value which puts cystine in the same class as other amino acids as far as absorption is concerned. The F-M method gives a lower value, in round nos., 30 mg. of cystine per 100 g. of body wt. per hr., in agreement with the work of Wilson (cf. *C. A.* 24, 3821), with the same method. If the value obtained by the F-M method is the true one, this method must be reacting to a desulfurized changed cystine left in the gastro-intestinal tract, a complex to which the other methods do not react. If this assumption is made that the excess of material reacting to the F-M method is only changed cystine, then the findings by the F-M method are the true rate of absorption of cystine by the white rat. Data are given to show that hydrolyzates of amino acids with no cystine present react pos. as cystine in the F-M method, but neg. in the other 2 methods, and that feeding glycine, and especially alanine, gives an increase in cystine value by the F-M method. Indications were obtained of an increase in glutathione content of the liver by feeding cystine. There are 27 references. J. A. KENNEDY

Modern conceptions on the antagonism of the sexual glands. ANGELO MIGLIA VACCA. *Boll. soc. ital. biol. sper.* 5, 1082-4 (1930). PETER MASUCCI

Relation between the calcium in the blood and the inhibitory action of the splanchnic nerve. E. GLEY AND H. JANKOWSKA. *J. physiol. path. gén.* 28, 789-99 (1930).—The inhibitory action of the splanchnic nerve on the movement of the intestines depends, as does its vasomotor action, upon the Ca content of the blood. When this content falls $\frac{1}{2}$, or $\frac{1}{3}$, there results a large diminution and sometimes even a temporary suppression of the inhibitory action. This action is reestablished rapidly under the influence of Ca injection. The inhibitory system is less sensitive to Ca deprivation than is the vasomotor system. L. A. MAYNARD

Suprarenal capsules and glutathione. A. BLANCHETIERRE, LEON BINET AND A. ARNAUDET. *J. physiol. path. gén.* 28, 818-21 (1930).—The suprarenal gland is the organ

which is richest in glutathione. The amt. of glutathione diminishes after ovariectomy and increases considerably during lactation. The venous blood of the suprarrenal is particularly rich in glutathione. Perfusing the glands with blood to which glutamic acid and cystine have been added causes an increase of glutathione in the perfusate and in the tissues of the gland, indicating that glutathione can be synthesized from these constituents. L. A. MAYNARD.

Psychic influences upon the calcium content of the blood serum. M. CH. PIERSTRÖM. *Acta med. scand.* 74, 378-95 (1911). The effect of induced sleep is the same as that of normal sleep upon the Ca level of the serum. Psychic functions which are characterized by a state of quiet produce an effect upon the chemistry of the blood which manifests itself in a lowering of the serum Ca, while psychic states characterized by restlessness have the opposite effect upon the serum Ca. S. MORRIS.

The ammonia content and ammonia formation in muscle and its relation to function and change in condition. VIII. The assumed participation of the amide nitrogen of blood and muscle protein in the chemical processes of active muscles. I. MANN. *Biochem. Z.* 231, 17-8 (1911). This is chiefly a critical examination of the reports of Bliss (C. 1, 23, 1910), who claims for the tissues the ability to synthesize from NH_3 ions and blood protein amide compounds and to utilize these amide groups for the neutralization of acids produced during work. It was found that neither in the human blood nor in the isolated frog muscle does amide N take part in the chemistry of the metabolic processes. IX. The position of ammonia formation in the series of chemical processes in active muscle. W. MOZDOLOWSKI, I. MANN AND C. LUTWAK. *Ibid.* 240-55. The curve of NH_3 formation during activity of muscle poisoned with CNHgCOH shows that the largest production of NH_3 takes place not during activity but during fatigue. A diminution of the creatine phosphide and the development of rigidity. S. MORRIS.

Experimental studies on the influence of air saturation on the resistance of red blood corpuscles to hypotonic salt solutions. GRANA NARGES. *Biochem. Z.* 231, 105-102 (1911).—Keeping animals in rarefied air greatly affects the resistance of their red blood cells to hypotonic NaCl solutions and after 6-8 days the min. resistance diminishes by 11.02-11.08% NaCl and the max. increases at least 0.08%. S. MORRIS.

Changes in the physical state of inorganic components of serum under reciprocal influences. HANA G. SCHMITZ. *Biochem. Z.* 231, 135-41 (1911).—Studies by means of ultrafiltration show that a colloidal Ca-P complex is formed in serum when either the Ca or P level is raised. A similar complex is also formed with Mg. The addition of CaCl_2 to serum increases the ultrafiltrability of the K and the addition of MgCl_2 increases that of Ca. The hypothesis of a colloidal K compound in serum is given support by these results. S. MORRIS.

Studies on blood glycolysis. III. Inhibition of glycolysis. H. E. HAMPHREY-SCHURPH AND KARL HRAUN. *Biochem. Z.* 231, 111-72 (1911) cf. C. A. 25, 1899. Under all conditions or agents which inhibit glycolysis of the erythrocytes (hemolysis, narcotics such as ether and CHCl_3 , diuretics, oxalates and the monophalogen derivatives of AcOH , $\text{CH}_3\text{XCO}_2\text{H}$) there is a splitting of inorg. P from org. combination, the chief source being the pyrophosphate fraction of the org. phosphate. With diuretics and $\text{CH}_3\text{XCO}_2\text{H}$ the inhibition of glycolysis is associated with a synthesis of difficultly hydrolyzable esters related to the hexosediphosphate fraction. From these considerations it is concluded that the pyrophosphate fraction is essential for the glycolytic process and is either the coenzyme or the activator of the coenzyme. This is substantiated by such facts as the following: the fermentation of yeast prepurified from, or poor in, coenzyme is somewhat activated by the addition of the pyrophosphate fraction from erythrocytes, and the inhibition of glycolysis by hemolysis, diuretics or $\text{CH}_3\text{XCO}_2\text{H}$ is accompanied by an accumulation of methylglyoxal from added hexosediphosphate. S. M.

Cholesterol in structural combination in the protoplasm. VI. Studies on the cow spleen. G. PERRYMAN. *Biochem. Z.* 231, 239-57 (1911), cf. C. A. 25, 1899. The spleen of the 1-yr. old cow contains 0.0916% total cholesterol per 100 g. fresh substance. Of this 81.1% are cholesterol + oxysterol (oxysterol = 0.1%) and 18.9% cholesterol esters. Of the latter the oleic acid of cholesterol was 7.5% and of oxysterol 24.0%. The spleen shows thus a higher cholesterol content than other organs which is associated with its phagocytic function of destroying red blood cells. The cholesterol set free in the process of erythrocyte destruction accumulates in the spleen. S. MORRIS.

Presence of ergosterol in the human brain. JAYNE H. PAGE AND WHITE MORGENTHAU. *Biochem. Z.* 231, 416-28 (1911).—The absorption spectra were studied of the purified sterol fraction from the brains of fetuses, newborn babies, lactating babies and adults. In the first three the absorption maxima of the brain sterols were at 291, 293 and

250 m μ resp. The absorption at 294 and 283 m μ belongs to ergosterol and its concn. in the cholesterol diminishes from 0.03% (fetus) to 0.01% in the 1 1/2-yr.-old baby. The substance corresponding to the absorption band at 250 m μ has not yet been identified. The ergosterol has been identified by means of digitonin pptn. and change in absorption spectrum under the influence of ultra violet radiation. The absolute amt. of ergosterol reaches a max. of 1 mg. in the brain of the 1-yr.-old baby, and then diminishes considerably. S. MORGULIS

The non-inhibition of cell respiration by hydrocyanic acid. OTTO WAABERG. *Biochem. Z.* 231, 493-7 (1931).—Dixon's discovery (*C. A.* 24, 426) that KCN affects only 60% of the cell respiration is shown to be based on error. Liver, kidney and spleen cells are injured by the phosphate buffer as used by Dixon. When these cells are experimented with in a Ringer NaHCO₃ + CO₂ medium their respiration is completely inhibited by KCN just as is that of yeast cells. Attention is called to several other expl. errors in Dixon's work. A no. of instances are cited demonstrating that injury of the cell frequently affects the inhibition of its respiratory activity under certain conditions. W. points out that the cardinal problem in cellular respiration is not whether O₂ or H₂ is activated but whether or not activation is catalyzed by Fe. From the fact that baker's yeast heated 90 min. at 52° practically loses the ability to oxidize physiol. substances, though it can still oxidize a mixt. of *p*-C₆H₄(NH₂)₂ and α -naphthol to indophenol, which is also inhibited by CO. Keilin (*C. A.* 25, 3719) concluded that the respiratory enzyme is the indophenoloxidase. The reaction with the heated yeast, however, is irregular and the explts. are interpreted as showing merely that the respiratory enzyme has been injured. S. MORGULIS

Skin and climacterium. CLARA NAEGELI AND MARIAN FELLNER. *Endokrinologie* 8, 81-121 (1931).—General. S. MORGULIS

The endocrine function of the placenta. ARTHUR A. PROBSTNER. *Endokrinologie* 8, 161-9 (1931).—From the fact that a patient who underwent an operation for double ovariectomy during pregnancy completed the term, gave birth to a normal child and still continued to secrete in the urine the sex hormone, just as the normal female does, it is concluded that the hormone must be produced by the placenta. S. MORGULIS

The mechanism of the action of the sex (follicular) hormone. PETER HAUPTSTEIN. *Endokrinologie* 8, 169-80 (1931).—Studies on mice show that the effectiveness of the follicular hormone depends only up to a certain degree upon the amt. of the hormone administered. Repeating the dose at short intervals may increase its effectiveness even a hundred fold. The loss of activity is rapid and independent of the dose and is apparently due to quick excretion. Following prolonged administration of nonphysiol. amts. of the hormone there is definite evidence of injury of the follicular app. of the mouse manifesting itself in inhibition of the follicular development and the failure of development of corpora lutea. The last fact is especially important in human therapeutics. S. MORGULIS

The chemical composition of animal fetuses. HANDE ØMEE. *Arch. Wiss. Prakt. Tierheilkunde* 62, 248-66 (1930).—Water, fat and ash were detd. in hog and sheep fetuses, and in hogs and sheep at various times shortly after birth. The water content of the fetus decreases, and the ash content and maternal free from fat and ash increase, as the fetus increases in size. The ash content of the fetus is less than that of the mature animal. W. GORDON ROSE

Partial diuresis. GIULIO LEGA. *Arch. farmacol. sper.* 51, 49-54 (1930).—When 1000 cc. physiol. NaCl soln. was given to normal fasting humans, the NaCl was promptly excreted, while an increased water elimination became noticeable after 4-5 hrs. Urea administered in physiol. NaCl soln. to healthy persons was immediately excreted, while the elimination of water and NaCl was retarded. Daily administration of NaCl to nephrotics did not at first exert an influence on the NaCl excretion, which, however, began to rise after a few days. Administration of 500 cc. water to nephritics with polyuria and with glomerular and vascular lesions caused a retarded diuresis, if any. Prolonged administration of urea to chronic nephritics with normal NaCl elimination and scanty urea excretion was followed in some cases by a prompt and prolonged urea excretion in the urine. In other cases with more severe lesions no urea was eliminated after administration of a single dose; in many of these cases, however, protracted urea administration was followed by increased elimination. No expl. data are given. G. SCHWOCH

Action of some organic and inorganic substances on the capacity for work of the gastrocnemius of the frog. VINCENZO PAMIANI. *Arch. farmacol. sper.* 51, 65-79 (1930).—One of the gastrocnemius muscles together with its nerve was isolated and immersed in a soln. of NaCl 0.60, KCl 0.03 and CaCl₂ 0.02% for 15 min. Then it was

put to work under isotonic conditions until it was completely exhausted. Meanwhile the substance to be tested was injected into the dorsal lymphatic sac of the frog. After 1 hr (15 hrs with glucose) the other gastrocnemius was isolated, immersed in the salt soln for 15 min and stimulated to complete exhaustion. From the muscular contractions as registered by the kymograph the work done was caled in g m and from the difference in work accomplished by the 2 muscles conclusions were drawn with regard to the favorable or unfavorable effect of the substance tested. Creatine, and to a slightly lesser degree NaH_2PO_4 , caused an increase in the capacity for work. K_2HPO_4 in very small doses acted similarly, while in large doses it had an injurious effect. The influence of creatine in combination with K_2HPO_4 or NaH_2PO_4 was not greater than that of one of the compds alone. Glucose alone or in combination with creatine or phosphates, or together with creatine and phosphates, always caused a distinct diminution in the work done. Other expts showed that within certain limits the isolated gastrocnemius of the frog had a greater capacity for work the longer it had been immersed in the salt soln. All conditions being equal, the right muscle had a greater capacity for work than the left muscle. A large bibliography is included. G. SCHWOCH

The participation of the nucleolus of hepatic cells in iron metabolism. GIOVANNI CORRADINI. *Arch farmacol sper* 51, 125-8(1930).—By means of the Berlin blue reaction C demonstrated that in the white mouse the nucleoli of hepatic cells participate in Fe metabolism. Normally the reaction is distinct only in the vicinity of the centrolubular vein, but after injection of colloidal Fe it was observed throughout the liver. Similar expts on rabbits and guinea pigs showed that in these animals the Berlin blue reaction could not be produced in the nucleoli of the hepatic cells. G. SCHWOCH

Effect of experimental hyperthyroidism on reproductive processes of female albino rats. CHARLES K. WEICHERT. *Physiol Zool* 3, 461-6(1930).—When normal female rats, which had shown regular estrous cycles, were fed with 0.25 g of desiccated thyroid daily, the diestrous condition persisted as long as the thyroid feeding was continued (6-22 days). The animals came into estrus 2-10 days after feeding was discontinued. The general health of these rats was not markedly altered, though the body wt had decreased approx 19% after 3 weeks of thyroid feeding. Sections of the ovaries revealed a no of corpora lutea along with follicles of all sizes. When pregnant mature rats were fed with desiccated thyroid (0.25 or 0.5 g daily) very severe reactions followed. Some of the animals resorbed their young, others died before or after delivery. The gestation period was prolonged in those rats which lived long enough to deliver their young. The young were born dead or died within 1 or 2 days. The animals appeared to be in good health until the last 2 or 3 days of pregnancy. The reactions were more severe in the groups receiving the larger dose. The possibility is discussed that the anterior lobe of the pituitary body may be involved. G. SCHWOCH

Separation of anterior-lobe substances and study of their individual effects. PEARL E. CLAUS. *Physiol Zool* 4, 36-57(1931); cf *Proc Soc Exptl Biol Med*, 27, 29-30—Armour's desiccated anterior lobe was extd with acid alc. In the attempt to sep the hormones the method of Fevold, Hisaw and Meyer (*Proc Soc Exptl Biol Med* 27, 604) was used with some modifications. Two substances were isolated. One of them (I) was a constituent of a cryst fraction; when injected subcutaneously it stimulated follicular development in adult female rats and induced early maturity in immature female mice and rats and in immature male mice. Identical crystals with the same physiol. properties were also obtained from urine of pregnant women, the epididymis of the bull, fish sperm, the thyroid gland and adrenal cortex. The other active fraction (II) was sol. in abs. alc. and contained a luteinizing hormone which was not obtained from a similarly prep'd ext of adrenal cortex or epididymis. The luteinizing fraction caused, in the ovary of an adult rat, the formation of atretic corpora lutea with the ovum in closed, but it did not induce early maturity. It also stimulated placentaloma formation, which does not occur with I. No increase in wt was noticed after adult rats had been treated with I or II for 6 weeks. Numerous references and 6 plates are given. G. SCHWOCH

The formation of bile pigments in tissue cultures. I. SÖMEGI AND M. CSABA. *Magyar Orvos Arch.* 31, 473-7(1930).—The spleens of chicken embryos and those of amphibians are able to produce bilirubin in tissue cultures from laked blood. The bile pigments were identified by the indirect diazo reaction. Undissolved hemoglobin (erythrocytes from the explantes) produces bilirubin in the same way but only in traces. The formation of bile pigments is a vital function of the spleen. H. T.

The conditions of metabolism which can permit the change of sex. PH. JOYET-LAVERGNE. *Compt rend* 192, 180-2(1931).—The changes of sex in animals, which have been described in recent years, are attributed to simple physiol. causes. the change

from male to female to the deposition of fat in the testicles, the change from female to male to an increase in metabolism, a decrease in reducing power of the tissues and changes in the composition of the blood. The last changes can be produced by tuberculous tumors (Crew, *Proc Roy Soc B* 95, 276, 1923) and are accompanied by the expected sex change. K. V. THIMANN

The relation between the effect of the estrus-producing hormone and a corpus luteum extract on the growth of the mammary gland. C. W. TURNER AND A. H. FRANK. *Science* 73, 211-6 (1931).—Development of the mammary glands characteristic of pregnancy could not be obtained in rabbits by injection of the estrus producing hormone alone. Injection of the hormone together with corpus luteum ext. prepared by Allen's method (*C. A.* 24, 4819) however, gave after 30 days a development of the mammary glands of castrated males similar to that of advanced pregnancy. The active principle of the corpus luteum ext. responsible for this effect was present in a water- and EtOH sol. ext. of the urine of pregnant cows, and this ext. could be substituted for the Allen's ext. in the above expts. K. V. THIMANN

Protoplasma-Monographien. Edited by FR. WENKA AND L. V. HEILBRUNN. Band V. PH. JOYET-LAVRAGNE. La physicochimie de la sexualité. Berlin: Gebrüder Borntraeger 457 pp. M. 32

G—PATHOLOGY

H. GIDEON WELLS

A modified method for the production of antipneumococcus serum in horses. EDWIN J. BANZHAF AND THEODORE J. CURPHEY. *Proc Soc Exptl Biol Med* 28, 279-80 (1930).—Horses were immunized by intramuscular injections of phenolized pneumococcus pleural exudate and intravenous injections of formalinized sediment of 18-hr. broth cultures. The sera of such horses had a greater therapeutic value than was indicated by their mouse protection unit content determined by comparison with antiserum produced by administering vaccine intravenously. C. V. BAILEY

Is ferratin precipitinogenic? LUDWIG HERTORN AND WM. H. WELKER. *Proc Soc Exptl Biol Med* 28, 293-4 (1930).—Ferratin, a deriv. of nucleoprotein high in Fe, is obtained by boiling the liver, filtering and pptg. with tartaric acid. It can be rendered practically free from blood and lymph protein. Hlog, beef and sheep ferratins were used in expts. on rabbits. The precipitinogenic power was slight and ferratin was not demonstrated to be a species specific antigen. C. V. BAILEY

Antigenic power of ultra-violet-irradiated tetanus toxin. EMERSON MCGRAIL AND HENRY WELCH. *Proc Soc Exptl Biol Med* 28, 494-5 (1931).—A mixt. of 2 tetanus toxins did to contain 1 m. l. d. per cc. was irradiated with the C arc at 25 cm. distance for 2 min.; the toxin was partly destroyed, half the inoculated guinea pigs developed late tetanus. A further irradiation of 2 min. rendered the material atoxic. By use of a quantity originally contg. 1 m. l. d., 9 guinea pigs were given 5 subcutaneous injections of this irradiated toxin at 6-7 day intervals and were inoculated 3 weeks later with freshly titrated tetanus toxin in doses of 1-10 m. l. d. Tetanus did not develop in the 45 days of observation. Ultra violet irradiation destroys the toxic but not the antigenic power of tetanus toxin. C. V. BAILEY

Absorption of hydrokollag from the obstructed bowel. H. A. CARLSON, H. J. DVORAK, F. W. LYNCH AND O. H. WANGENSTEEN. *Proc Soc Exptl Biol Med* 28, 542-4 (1931).—Hydrokollag, a colloidal suspension of graphite, was not absorbed from the obstructed intestine of the dog or of the rabbit. It was absorbed from the peritoneal cavity and possibly from the ulcerated intestine. C. V. BAILEY

Gastric acidity in diabetes mellitus. Its chemical significance based on a study of one hundred cases. I. M. RABINOWITCH, A. F. FOWLER AND B. A. WATSON. *Arch Internal Med* 47, 384-90 (1931).—A study of normal cases shows a possible incidence of 20% who have achlorhydria. This incidence is, however, only one-half that found in diabetes mellitus, suggesting that this disease is accompanied by low gastric acidity. The etiology of this condition is discussed. J. B. BROWN

The metabolism of galactose. I. Conditions underlying the use of galactose in tests on the function of the liver. HARRY SHAY, EUGENE M. SCHLOSS AND MILTON A. BELL. *Arch Internal Med* 47, 391-402 (1931).—Galactose is very suitable for testing liver function for the following reasons: readily obtainable pure, readily absorbed from the intestine, convertible with difficulty by the liver into glycogen; not utilized by other tissues than the liver, in the general circulation it is readily excreted in the urine, regardless of the condition of the kidney or endocrine activity. J. B. BROWN

The blood calcium in diabetic retinits. R. D. LAWRENCE, KATE MADDERS AND

H. R. MILLAR. *Brit Med J.* 1930, II, 559.—The observation that diabetic retinitis is accompanied by low blood Ca is not confirmed. J. B. BROWN

The relationship of blood uric acid content to the state of renal function in nephritis. CHRISTOPHER JOHNSTON. *J Clin Investigation* 9, 555-9(1931).—Simultaneous detn. of blood uric acid and urea clearance showed the former to be of little value in the estn. of kidney function. Despite 80% loss of urea-excreting power by the kidney, uric acid may be normal. J. B. BROWN

The excretion of intravenously injected bilirubin as a test of liver function. GEORGE A. HARROP AND E. S. GUZMAN BARROW. *J Clin Investigation* 9, 577-87(1931).—The bilirubin excretory power is the most delicate method for testing liver function. J. B. BROWN

Studies of serum electrolytes. VII. The total base and protein components of the serum during lobar pneumonia with a note on the gastric secretion. F. WM. SUNDERMAN. *J Clin Investigation* 9, 615-33(1931).—Study of the base and protein constituents of the serum in patients with lobar pneumonia showed a decreased concn of Na in proportion to the decrease in concn of total base. In over half of the detns an increase of K appeared before and after the crisis. During the febrile period Ca was decreased, while the Mg values were variable. A decrease in albumin was balanced by increase in globulin, making the total protein concn nearly normal. Protein-bound base usually fell during the febrile period. Gastric acidity was low until after the crisis. There was no apparent correlation between serum total base or chloride with gastric acidity. J. B. BROWN

The cause of death in liver autolysis. EDMUND ANDREWS. *Surgery, Gynecol Obstet* 52, 61-6(1931).—Implantation of sterile liver (in vivo autolysis) produces a toxic reaction, described as autolytic peritonitis. The toxic agent probably falls within the albumose group. J. B. BROWN

Action of parathyroid graft on calcemia in normal or parathyroidectomized dogs. J. T. LEWIS AND REBECA GEASCHMAN. *Compt rend soc biol* 103, 1281-3(1930).—See C. A. 25, 337. E. J. C.

Leucine and tyrosine in urine in lung tumors. HANS ASPERGER. *Wiener klin Wochschr.* 43, 1281-4(1930).—Leucine and tyrosine are excreted in the urine in increasing quantities with the progress of carcinoma of the lungs. Their presence in the urine is of diagnostic value in doubtful cases. D. B. DILL

The influence of ergotamine and liver diet on liver function. RICHARD BAUER AND OSKAR WOZASEK. *Wiener klin Wochschr.* 43, 1337-41(1930).—The administration of 40 g. of galactose results in galactosuria in some cases of liver disease when on a liver diet but does not occur with normal subjects on a liver diet. D. B. DILL

Distribution of electrolytes in serum and cerebrospinal fluid. I. Calcium and magnesium. ZDENKO STARY, ADALBERT KRAL AND RUDOLF WINTERITZ. *Z. ges. expil. Med* 66, 671-91(1929), cf. C. A. 24, 3543.—The Ca content of cerebrospinal fluid approaches that of the dialyzable fraction of the serum. Dialysis of serum against spinal fluid produces only a slight change in the Ca distribution. The Ca content of cerebrospinal fluid is more constant than the serum Ca in the same individual, while Mg is the reverse. The spinal fluid Mg is higher than the serum Mg. A series of compensation dialysis experiments suggests that the serum Mg is in a complex anion. The ratio of Mg in spinal fluid and serum varied between 110 and 200% for non meningitic cases. Tuberculous meningitis had a ratio less than 100%. II. Potassium and sodium. *Ibid* 691-701.—The K content of cerebrospinal fluid in a group of miscellaneous cases averaged 11.7 mg. % with an av. of 21.7 mg. % for the serum, giving a distribution coeff. of 53.92%; similarly for Na 235 mg. % for spinal fluid and 321 mg. % for serum, and a quotient of 91%. Dialysis of serum and cerebrospinal fluid results in both fluids having the same K concn. There were no characteristic changes in the distribution ratio of K or Na in the various diseases studied. F. L. DUNN

Behavior of free cholesterol and its ester in the blood and organs of the rabbit following artificial obstruction of the bile. E. VON ROSZTOCZY. *Z. ges. expil. Med* 68, 690-700(1929).—The rise in free cholesterol occurs in spite of the defective absorption of cholesterol resulting from the failure of the bile acids to be secreted because of the obstruction. The ester cholesterol showed a decrease in the organs while it was increased in the blood. The liver damage resulting from biliary obstruction is a factor in determining the relative amounts of free cholesterol and its ester. Part of the increase in cholesterol is due to the erythrocyte destruction accompanying biliary obstruction. The bibliography contains 75 references. F. L. DUNN

The specific dynamic action of protein in disturbed endocrine function. M. SEREISKII AND SOPHIE JISLIN. *Z. ges. expil. Med* 69, 321-36(1930).—Disease of the

thyroid resulted in a reduction or absence of the sp dynamic action for protein. This change in hyperthyroidism was due to the satn. of the body with thyroid substance and to the over stimulation of the vegetative nervous system. In hypothyroidism the lowered sp dynamic action is a reflection of the general torpidity of bodily processes. The metabolism in thyroid disease is essentially endothermic in type. F. L. DUNN.

Serum proteins in secondary syphilis. A. SEZARY AND P. MARTINET. *Bull Mem Soc Med Hopit Paris* 1930, 945-7. *Bull Hyg* 6, 55.—The authors investigated the relation of albumin and globulin in the sera of 6 syphilitics in the secondary stage showing no nephrosis. The normal ratio is 16/1. The authors' figures range from 0.67 to 0.97/1. They assume the change due to changes in the metabolism brought about by spirochetal septicemia, not to nephrosis.

GEORGE R. GREENBANK

Present status of the investigation of the cause, and of the geographical distribution, of mottled enamel, including a complete bibliography of mottled enamel. FREDERICK S. MCKAY. *J Dental Research* 10, 561-8 (1930).—“The evidence has been so overwhelmingly conclusive, to the effect that the productive influence resides in the drinking water, that the investigation now rests upon this hypothesis.” The bibliography contains 47 references.

JOSEPH S. HEPBURN

Cellular individuality in the higher animals with special reference to the individuality of the red blood corpuscle. II. CHARLES TODD. *Proc. Roy Soc (London)* B107, 197-205 (1930).—Examination of a family of Plymouth Rock chickens by means of exhausted immune iso-agglutinating sera revealed that the corpuscles of no 2 chickens were exactly alike, the cells of different individuals differed in immunological behavior from a close resemblance to a very marked contrast. Individual chickens, immunized with the blood of their own brothers and sisters, yielded active, highly sp agglutinating sera. The ease of formation of the agglutinins varied with the degree of difference in character between the erythrocytes of the injected chicken and those used for immunization. “The red blood corpuscle must be regarded as a ‘multiple antigen’ in the sense that it contains a large number of different antigenic units or ‘receptors,’ which apparently behave as independent units when hereditarily transmitted.”

JOSEPH S. HEPBURN

Heat of rigor of mammalian muscle. E. C. SMITH. *Proc Roy Soc (London)* B107, 214-22 (1930).—Neither the disappearance of glycogen nor the formation of lactic acid suffices to explain the stiffening in rigor. A third chem. change, as yet not clearly understood, must also be involved.

JOSEPH S. HEPBURN

Lactic acid content of the blood in several diseases, especially malignancy. E. M. GELSTEIN AND M. I. FRANKSTEIN. *Z klin Med* 111, 563-81 (1929).—The av lactic acid content of the blood in healthy well nourished persons is 9.08 mg per 100 cc. It was increased in anemia, leucemia and cardiac disease with cyanosis and decreased oxidation in the tissues, and also in 50% of the cancer patients studied. A dose of 100 g of dextrose or sucrose produced a temporary increase in the lactic acid content of the peripheral blood, the av concn being 66.9 mg per 100 cc. at the end of 1 hr. This phenomenon was still more marked in some cancer patients.

JOSEPH S. HEPBURN

Precipitation of sediments and the formation of calculi in the urinary tract. JOHANNES MEYER. *Z klin Med* 111, 613-87 (1929).—Pure uric acid calculi form in a normal, strongly acid urine with pH approx 5.0. Mixed calculi, composed of pure, quite dense uric acid, urate, oxalate and Ca phosphate layers, arise in normal urine of pH 6.0. Pure Ca phosphate calculi form in normal neutral urine of pH approx 7.0. Hard, stratified calculi, composed chiefly of $MgNH_4PO_4$, Ca phosphate and NH_4 urate, form slowly in an ammoniacal urine with a faintly alk reaction ($8 > pH > 7$). If layers contg considerable amts of $CaCO_3$ be present in these calculi, then the reaction at their formation has been decidedly more alk. Soft free calculi, consisting chiefly of $MgNH_4PO_4$, $CaCO_3$ and NH_4 urate, form rapidly in a strongly alk. urine with $pH > 8$. Org compds are simultaneously pptd, but the degree of satn. of the urine with crystal line compds decides when the calculus grows, and the compn of the deposited layers. An elaborate bibliography is appended.

JOSEPH S. HEPBURN

Kidney function tests and blood-urea determination. H. KRAUCKENBERG. *Arch Gynakol* 143, 56-71 (1930).—The various methods for detg the functional activity of the kidney are reviewed. Most of them are too complicated for the average practitioner and require elaborate lab equipment and often special prepn of the patient. The detn. of the urea content of the blood by means of the azotometer is suggested as a reliable method for the average practitioner as the test is easily made and requires no special prepn of the patient.

HARRIET F. HOLMES

The xanthoproteic reaction in blood from which protein has been removed in pregnancy. A. H. EYINGER. *Arch Gynakol* 143, 368-75 (1930).—The xanthoproteic reaction according to the method of Becher was carried out on blood from which the

protein had been removed from 841 pregnant women, 10 cases *pre partum* and 175 cases *post partum*. The xanthoproteic reaction was neg. during pregnancy and *pre partum*. *Post partum* a positive reaction was obtained in 15%; and the reaction was obtained most often between the 5th and 10th days *post partum*. The reaction of the total xanthoproteic content makes it seem probable that an increase of the intermediary ammonia and amino acids is an indication of anabolic processes in the uterus in the puerperium. A pos. xanthoproteic reaction was obtained in 40% of the toxoses of pregnancy, nephropathy and eclampsia. The xanthoproteic reaction is often markedly increased in cases of carcinoma of the uterus after a ray or radium irradiation. Possibly the reaction will furnish a means for following the effect of irradiation in carcinoma cases.

HARRIS I. HOLMES

Carbohydrate metabolism in cancer of the uterus after Röntgen and radium irradiation. I. LUCAS. *Arch. Gynäk. 143, 380* (1930). In 10 cases of carcinoma of the uterus after treatment with x-ray or radium irradiation the sugar content and lactic acid content of the blood remain normal with the exception of a transitory increase in lactic acid content shortly after irradiation. The values for the alkali reserve indicate a hypocalcemia of the blood. There was a decrease of alkali reserve 2-4 hrs. after the irradiation which was of short duration and 24 hrs. later there was an increase of alkali reserve which was slightly increased in each further irradiation. However there is never a return to normal CO₂ binding power. The transitory increase in lactic acid shortly after irradiation coincides in time with the decrease in alkali reserve. The products of carbohydrate metabolism furnish no explanation for the persisting hypocalcemia in patients with carcinoma of the uterus.

HARRIS I. HOLMES

The Manielloff reaction for pregnancy. I. FARRWELL AND K. HILL. *Arch. Gynäk. 143, 470* (1930). The Manielloff reaction (C. A. 24, 1315) for pregnancy was tried on sera from pregnant women from healthy nonpregnant women and healthy men, and from patients with conditions that frequently give pos. results with other serological tests for pregnancy. The reaction was 81% pos. in the later months of pregnancy, but almost uniformly neg. in the earlier months of pregnancy. The reaction, therefore, is of no practical importance for the diagnosis of pregnancy. The reaction is not sp.

HARRIS I. HOLMES

The two types of bilirubin diast reactions in serum, with a hypothesis on the nature of the bilirubin in the serum from hemolytic jaundice. G. HENRIK. *Brit. J. Exptl. Path. 11, 115* (1930). An attempt has been made to set the probable nature of the 2 kinds of diast reaction for bilirubin obtainable in icteric sera. The so-called direct reaction, as it occurs in bile or in the serum of patients with obstructive jaundice, is most probably due to the presence of Na H bilirubinate. The so-called indirect reaction as it occurs in hemorrhagic exudates and in the serum of patients with hemolytic jaundice is probably due to the presence of free bilirubin in an unsplit solvent which is not in bile. The bilirubin solvent (it may be a lipid or steroid) need be present in only relatively small amounts to retain the bilirubin in solution on the addition of alc. The bilirubin may be in combination with its solvent or some other serum constituent so that it prevents salt formation and oxidation. It would not be dialyzable or extractable with CHCl₃. It thinks such a hypothesis to account for the delayed type of bilirubin would fit the findings better than the assumption of the presence of colloidal bilirubin.

H. I. H.

Phosphorus, total calcium and diffusible calcium content of the blood sera of lepers and their relation to bone changes. JEROME D. WIDNEY AND MILVARY RICE. *U. S. Pub. Health Rept. 46, 611* (1931).—Sera from 47 lepers were analyzed for Ca and inorg. P; the results were found to be within the range usually considered normal. Sera from 15 normal, healthy young men were examined for inorg. P, total Ca, diffusible Ca; the % of Ca that was diffusible, the Ca/P and the diffusible Ca/P balance ratios were determined. Sera from 51 lepers, representing the various types and stages of progression of the disease, were similarly examined. The diffusible Ca and the % of total Ca that was diffusible averaged considerably lower in the lepers than in the normal young men, and in only a few instances did the diffusible Ca in the lepers approx. the level found in the controls. Radiographic studies of the bones of the hands and feet of 18 of these lepers showed bone atrophy in 41 (91.0%), the condition manifesting itself either as decalcification or absorption. The writers appreciate that the records here submitted do not include a sufficient number of cases, observed over a sufficiently long period, to warrant the drawing of conclusions. It is thought, however, that deviation from the presumed normal has been marked enough to justify confirmation of the investigation.

J. A. KENNEDY

Antigenic value of scarlet-fever streptococcus toxin modified by the action of formalin. M. V. VIKTOR. *U. S. Pub. Health Rept. 46, 601* (1931).—A tulicly

reviews the data presented by various independent workers. The exptl work here presented confirms the findings of others that scarlet fever streptococcus toxin can be rendered much less toxic by subjecting it to heat and formalin. The antigenic value of the toxin apparently is not destroyed either by the action of the formalin or by the prolonged exposure to 37° to 38°. The amt of antigen tolerated by susceptible individuals in each injection is greatly increased by this method of detoxification, so that the number of doses required for active immunization may be reduced. At the same time the reactions following the injections occur much less frequently and their severity is markedly diminished. Further experimentation may, it is hoped, develop a technique for removing the remaining toxic factor, thereby eliminating all disagreeable reactions and permitting the use of greater quantities of antigen. J A KENNEDY

The reactions of immunity among invertebrates. J CANTACUZÈNE *Arch roum path expul microbiol* 1, 1-75(1928).—The introduction into the celom of *Sipunculus nudus* of sheep red corpuscles causes a certain no. of reactions on the part of the organism which have as their final effect the elimination of foreign particles in suspension in the cavity liquid, their reabsorption and the ultimate incorporation of the residue of this process into the brown bodies of the general cavity. The reabsorption is brought about solely by the phagocytes, the phagocytic action is facilitated by the agglutinant action of the viscous secretions. The process is characterized by an overproduction of all the cellular elements, and an intensification of their secretions, as well as a marked tendency toward agglutination and pptn. MARIE MAXIM

Investigations in the resistance of diphtheria and dysenteric toxin to different concentrations of hydrogen ions. C IONESCO MIIHAIEȘTI AND A DAMBOVICĂNU *Arch roum path expul microbiol* 1, 115-21(1928).—The destruction of the diphtheric toxin by the different concns in H or OH ions is not immediate. For zone pH 4.7-4.9 and 9.6-10.1 the toxicity is almost unchanged for 30 min. For 24 hrs the limits are pH 5.7-5.9 and 9-9.4. Likewise for dysenteric toxin the toxicity does not diminish until after a certain time, 30 min. in a medium of pH 2.07-2.15 or pH 11 does not suffice for any change, 24 hrs. of pH 2.5-3 completes the action of the toxin. MARIE MAXIM

Appearance of a desensitizing and hypotensive substance in the organism after blocking of the reticulo-endothelial system. I MODOVAN *Arch roum path expul microbiol* 1, 167-77(1928).—The intravascular or intraperitoneal injection of substances provoking the blocking of the reticulo-endothelial tissue brings about in the serum of inoculated animals the appearance of a desensitizing hypotensive, thermo-resistant substance, sol. to acids, alc., water, easily dialyzable through membranes of collodion and which has not been identified. MARIE MAXIM

Investigations on blood lipase in different forms of tuberculosis. I NICOLAU AND O ANTINESCU *Arch roum path expul microbiol* 1, 437-51(1928).—Fat, detd. according to the stalagmometric method, is generally diminished in the blood of tuberculous patients, its variations are in relation to the general state of the sufferer and not to the degree of tuberculous lesion. MARIE MAXIM

The glycogenesis of the sarcoma of Peyton Rous. E C CRACIUN, *Arch roum path expul microbiol* 2, 105-11(1929).—The sarcoma of Rous presents a quite abundant and irregular glycogenesis at the level of coagulation necrosis as well as with liquefying necrosis. There is no relation between the glycogenesis and the virulence of this tumor. MARIE MAXIM

Investigations on specific substances (residual antigens) of the anthrax bacilli. D COMBIESCO, E SORU AND S STAMATESCO *Arch roum path expul microbiol* 2, 291-312(1929).—A water sol. substance having the properties of polysaccharides has been sep'd from *B. anthracis*. This substance is not toxic for white mice and guinea pigs, does not form antibodies and gives a pptn. with specific serums. The rotary power after hydrolysis is -20 , it contains hexoses and pentoses. MARIE MAXIM

Glycogenesis and regeneration of voluntary muscles. E C CRACIUN *Arch roum path expul microbiol* 2, 313-23(1929).—The sarcoblasts and granulation tissue contain no glycogen, this substance is often present in abscesses. M M

The heliotropic property of cholesterol in relation to cancer of the skin. A II ROFFO AND FR PILAR *J physiol path gen* 28, 854-6(1930).—Exposure of the skin to sunlight or ultra violet light increases its cholesterol content. In the authors' experience 92% of the skin tumors of the lace occur on parts directly exposed to the sun. They conclude that there is a relationship between the heliotropic properties of cholesterol, its fixation in the skin and tumor formation. L A MAYNARD

The chemical investigation of lung tissue in pneumoconiosis. LUIS ROSSI AND SANTIAGO A CELSI *Semana méd* (Buenos Aires) 1931, I, 629-34.—In the lungs of

minerals in Dollyda, SiO_2 , P_2O_5 , Al and Sn were found, the last at a ratio of 1:199610% calculated on dry or 0.272% calculated on fresh substance. A. R. MURPHY.

Studies on the subfractions of the serum globulins and albumins. II. The distribution of lipids, preprecipitogens and bacterial agglutinins among the separate fractions of serum. It LURIO AND R. KATZ. *Biochem. Z.* 231, 39-44 (1934).—The largest lipid content is found together with the albumin and the smallest with the englobulin fractions of serum. Among the subfractions of englobulins and pseudoglobulins the highest total lipid and lecithin content is found in the proteolysed in NaCl and the smallest in the water and ones. The largest amount of the sp. preprecipitogens is also in the NaCl and englobulin and pseudoglobulin and the least in the albumin fractions I and II. Normal agglutinins are chiefly associated with the englobulin and are practically absent in the albumins. Among the globulins the NaCl and fraction has the largest amount of agglutinin while the NaCl and fraction has the smallest amount. III. Elementary composition, distribution of amide nitrogen, humin nitrogen, diamino nitrogen, mono- or non-amino nitrogen, of tryptophan, and of carbohydrates in the separate subfractions of cow serum. It LURIO AND P. HAAS. *Ibid.* 472-482 (1934). The water absorption (% increase in wt. of dry powder suspended in 100) of the different serum protein fractions is as follows: water sol. englobulin 88.12%, NaCl and englobulin 96.13%, NaCl and pseudoglobulin 83.10%, and NaOH and englobulin 84.12%, water and pseudoglobulin 88.11%, NaCl and pseudoglobulin 81.12%, NaCl and pseudoglobulin 88.05%, albumin I 10.6.11%, albumin II 0.6.12.5% and albumin III 8.0.12.8%. The elementary composition of the different fractions in the same order, are: % C 61.80, 62.14, 66.76, 49.61. % H 10.2, 11.19, 17.14. % N 16.25, 16.00, 17.83. % O 13.00, 16.45, 13.25, 16.13, 14.99, 16.02. % S 1.61, 1.65, 1.71. % P 1.67, 1.61, 1.31. % Cl 1.8, 2.37, 2.18, 2.27. % I 1.110, trace, 0.16, 0.21, 0.12, 0.29. 0.00. 0.07, 0.03, 0.15. % CH₃ 0.0, 0.0, 0.06, 0.40, 0.47, 0.05, 0.42, 0.11, 0.67, 0.00, 1.01. % CH₂ 0.0, 0.10, 0.67, 0.18, 0.11, 0.83, 0.68, 0.04, 0.69, 0.06. The percentages of amide N in the different fractions were 4.5, 1.6, 3.2, —, 4.3, 1.6, —, 4.8, 1.1 of humin N, 4.8, 2.1, 2.2, —, 1.2, 1.1, —, 1.6, 1.7, —, 1.4 diamino N, 27.6, 26.8, 27.0, —, 21.8, 29.0, —, 32.6, 32.7, —, of mono- and non-amino N, 18.2, 71.8, 61.8, —, 72.2, 69.1, —, 33.8, 62.6, —. The average percentages of carbohydrate were 0.81, 0.99, 2.79, 8.70, 0.98, 0.51, 0.40, 1.17, 0.65, 0.65. H. MURPHY.

Studies on antileish. PAUL KREMER. *Biochem. Z.* 231, 178-182 (1934).—The antileish is adsorbed with kaolin from an antiserum. The relation of the antileish in the serum proteins is thereby not affected since the latter are likewise adsorbed. The adsorption of both being proportional to the quantity of kaolin. The adsorption is complete in 10 min., neither the order of mixing nor the degree of dilution of the serum being of any consequence. Small amounts of kaolin adsorb the antileish and proteins without reacting NH_2 groups, whereas with larger amounts proteins with reacting NH_2 groups also appear. A Mg phosphate buffer of pH 7.17 leaches out the antileish only to a small extent, but this is improved by the addition of normal serum to the buffer. NaCl does not leach out the antileish even in the presence of a great excess of serum. Infusorial earth Cy is an excellent adsorbent, and the antileish can be very well leached out with phosphate buffer (and with NaCl), the ease depending upon the concentration of the buffer. Probably the PO_4 ions are essential for the leaching out. With increasing acidity of the buffer solution it becomes less suited for leaching out purposes. Furthermore, the PO_4 ions also tend to displace the proteins and antileish from the adsorbent, since the latter treated with phosphate loses some of its adsorbing power. The leaching out was repeated 3 times, most of the antileish together with most of the protein being removed after the first leaching, each successive leaching, however, removing a constant portion of the antileish present in the adsorbent. The second leaching yields antileish poorer in protein. Through fractional pptn. with $(\text{NH}_4)_2\text{SO}_4$, much protein can be removed without affecting the antileish content. Further purification can be effected through electrodialysis, etc. Antileish is not attacked by pepsin, carboxypeptidase or trypsin. H. MURPHY.

Rose bengal in the examination of the liver function. ALBERT REICHERT. *Arch. Pharmacol.* 51, 25-31 (1934).—It concludes from the analysis of liver-function tests on 12 patients that with the Rose bengal test the eliminating power of the liver, but not the entire functional ability of the hepatic cells, is tested. When bile pigment or urobilin is found in the urine and when the Rose bengal test is positive, the functional ability of the liver is altered. The reverse is not true, however, since the elimination may be normal, though the more complex functions of the liver may be impaired. G. BENJAMIN.

The gastric secretion in fasting patients. FERNANDO REICHERT. *Arch. Pharmacol.*

per. 51, 55-63(1930) —In 70 patients with gastric disturbances the gastric secretion was studied with respect to quantity of fluid, free HCl and total HCl. The samples were taken from patients in the fasting state or after a meal. In the patients in the fasting state a distinction is made between the "immediate secretion" which represents the 1st sample taken and the "late secretion" comprising samples taken after the removal of the 1st sample, in 10-min. intervals for the following 30 min. An "immediate fasting secretion" was found in 94% of the cases. In 72% the quantity of juice was less than 40-45 cc. Quantities in excess of 40 cc. were found in diseases associated with hypersecretion. Quantities larger than 60 cc. contained food residues and were due to pyloric stenosis. Free HCl was found in 78.57% of the cases, this percentage increases with the "late secretion" or after a meal. In general the immediate fasting secretion varied with the different gastric disorders and a relation between secretion and disease could be established. In diseases associated with hyperchlorhydria the secretion was generally abundant (25-60 cc.), while in diseases with normal or subnormal HCl the secretion was small (5-25 cc.). In the diseases with little secretion the content of free HCl was very slight. About 30 cc. is considered as the normal immediate fasting secretion. The "late secretion" was in general more abundant in cases with hyperchlorhydria and less abundant in cases with hypochlorhydria and anachlorhydria, the av. quantity secreted within 30 min. is 35-50 cc. The content of free HCl was generally greater in the late secretion than in the immediate secretion and slightly smaller than in the secretion after the meal. The late secretion was quite frequently (30%) contaminated because of duodenal regurgitation, therefore this test is only of relative value. The importance of the examn. of the fasting gastric secretion is emphasized with respect to its value in the diagnosis. G. SCHWOCK.

Uremia in diabetes mellitus. ALFREDO RICCI *Arch. farmacol. per.* 51, 41-5(1930) —The uric acid of the blood was detd. in 14 fasting diabetic patients who in the days previous to the expts. had been given a diet low in purine-contg. foods. The uric acid was normal in the cases treated with insulin and was increased in the cases not treated with insulin. Since the detns. were carried out according to the method of Lolin and Denn, and since xanthine and guanine give the same reaction with this method, it assumes that the oxidation of the intermediate products in the purine metabolism, for which the kidneys are impermeable, is retarded, thus giving rise to the accumulation of these products in the blood. This hypothesis is supported by expts. which were conducted to exam. the degradation of the exogenous purine substances. One normal person and 11 diabetics were given a meal consisting of 150 g. liver and 50 g. calf's tripe. Three hrs. later the uric acid was elevated in the healthy person and 6 hrs. later it had almost returned to its initial value. In the diabetics the uric acid was increased after 1 hrs. and was still far above normal after 6 hrs. The uric acid curve of the diabetic treated with insulin was similar to that found in the normal person. These results indicate that the purine metabolism is intimately linked with the carbohydrate metabolism, when the latter is deranged, as in the diabetic, the purine metabolism is also affected. Insulin administration to diabetics restores the carbohydrate metabolism to normal and consequently the purine metabolism is also restored to normal. G. S.

Chloremia and azotemia in cardiopathy. SERAFINO MANCA *Arch. farmacol. per.* 51, 81-94(1930) —Urea N and Cl of the blood and Cl of the urine were detd. in 20 patients with heart disease who had been kept on a const. diet for 2 or 3 days before the test. The patients with compensated heart disease showed values similar to those obtained from healthy people. In the patients with non-compensated heart disease the Cl of the blood and urine was diminished while the non-protein N of the blood was augmented. The more the non-compensated state was reduced by proper treatment, the more the Cl and urea N approached normal values, without ever completely equaling them. These findings are probably related to the fact that in these cases slight renal lesions persist even after compensation has occurred. G. SCHWOCK.

Lactacidemia in renal insufficiency. ALFREDO RICCI *Arch. farmacol. per.* 51, 141-4(1930) —Lactic acid was absent in the plasma of 12 patients with partial renal insufficiency, it was present (0.019 and 0.015%) in 2 cases of complete renal insufficiency. Its presence signifies an unfavorable prognosis. G. SCHWOCK.

The question of the effect of unilateral eye extirpation. BERNHARD DÖRKEN *Biol. generalis* 6, 511-52(1930) —D. had found in the course of earlier investigations on frog larvae that after unilateral eye extirpation the normal development of the midbrain extremities and non-operated eye is inhibited. The controls developed normally in the same water. Further expts. revealed that the water used in those expts. contained a small amt. of Zn which originated from a zinc plated storage basin. Recent investigations showed that in many cases the development of extremities, midbrain and eye, was

that of normal fowls. This is a primary change resulting apparently from an overflow into the blood stream of the glycolytic enzyme, which Warburg showed to be in tumor tissue in greatly increased amounts. An examn of the glycolytic activity of the blood of 16 patients with malignant tumors gave values that are within normal limits. The detn of the glycolytic activity of the blood is, therefore, of no clinical value in studying malignant conditions. The discrepancy between the glycolytic activity of the blood of hens with the Rous sarcoma No. 1 and that of the blood of patients with malignant conditions is probably due to the greater ratio of tumor tissue to normal tissue in the fowls with the sarcoma. **I. G. WELLS**

Immunological relationships among the pneumococci. V. Anaphylaxis and precipitation between antigens and antisera of yeast of type II pneumococci. **JOHN V. SLOGG AND JAMES M. NEILL. J. Exptl. Med.** 53, 527-34 (1931).—A report is given of the interactions of anaphylaxis and pptn. between antigens and antisera derived from Type II variety of pneumococcus and from one variety of yeast. That the reactions occurred only with Type II and not with Types I and III is proof that the pneumococcus antigen responsible for the anaphylaxis of the antiyeast sensitized animals was the type-sp. carbohydrate (S). **C. J. WEST**

Bilirubinemia. E. S. GUZMAN BARRON. *Medicine* 10, 77-133 (1931).—An exhaustive critical review of the literature of the physiology and chemistry of bilirubin, methods of detn. and their clinical application, with a bibliography of 200 references. **I. G. WELLS**

Biochemistry of Rous sarcoma of fowls. L. Biophysics. V. Coagulation of the serum. **ITIKI HOLZ. Z. Krebsforsch.** 33, 133-45 (1930), cf. *C. A.* 24, 5858.—In all cases the quantity of serum proteins of normal fowls showed wide variations, quite as great as in disease, and the proteins show the same coagulation properties, whether normal or pathol., with most precipitants. With CuSO_4 and $\text{Cd}(\text{NO}_3)_2$, some irregular differences were obtained. VI. Alterations in the crystal form of sodium chloride. *Ibid.* 140-54.—A difference was found in the crystal forms when serum from normal and sarcomatous fowls was evaporated to dryness. **I. G. WELLS**

Vitamin A content of melanosisarcoma of horses. **S. V. GUDJONSSON. Z. Krebsforsch.** 33, 155-7 (1930).—The amt. present in the tumors and in the adjacent muscle tissue is very small. **I. G. WELLS**

Sulfur impoverishment of the body and the tendency to cancer. **GRUMBE. Z. Krebsforsch.** 33, 376-7 (1931).—Purely speculative. **I. G. WELLS**

The stimulation of epithelial growth with synchronous disturbance of connective tissue growth in tissue cultures by substances which affect surface tension. **M. KATZENSTEIN AND ELSE KNAKE. Z. Krebsforsch.** 33, 378-410 (1931).—Substances with high surface effects have a sp. action on tissue cultures of the chicken embryo. Cholesterol or amyl alc. depresses connective tissue growth and stimulates epithelial proliferation. Other substances exert a similar effect in direct proportion to their surface effects. **I. G. WELLS**

The American Journal of Clinical Pathology (*New journal*). Published bimonthly by the American Society of Clinical Pathologists at Mount Royal and Guilford Aves., Baltimore, Md. Vol. 1, No. 1 appeared Jan., 1931. \$3 per year.

H—PHARMACOLOGY

A. N. RICHARDS

Avertin. F. B. PARSONS. *Brit. Med. J.* 1930, II, 554-7, cf. *C. A.* 24, 5373.—A review of the chem. properties and the clinical uses of this drug. **J. B. BROWN**

Sodium amytal and other derivatives of barbituric acid. L. G. ZERFAS. *Brit. Med. J.* 1930, II, 897-902.—A review. **J. B. BROWN**

The elimination of ethyl iodide after inhalation and its relation to therapeutic administration. **HERRMAN L. BLUMGART, DOROTHY ROUCKE GILLIGAN AND JACOB H. SWARTZ. J. Clin. Investigation** 9, 635-50 (1931).—Following the introduction and inhalation of 1 to 3 cc. of EtI in the inspired air, 45% is immediately exhaled during the treatment, 9% is exhaled during the next 2 hrs. and 31% is excreted by the kidneys as iodide. A total of 85% of the dose is accounted for, the remainder being probably excreted in the sweat, etc. Eighty-five % of the urinary excretion appears in the first 24 hrs. and 97% in 48 hrs. To maintain high iodide concn in the body daily administration must be given, the treatment being omitted every 3rd or 4th day to avoid iodism. The % of EtI retained in the body is independent of the size of the dose. **J. B. BROWN**

Studies in congestive heart failure. VIII. The effect of the administration of

diabetic potassium phosphate on the potassium content of certain tissues. J. ALFRED CALHOUN, GLENN E. CULLEN, GURNEY CLARKE AND T. R. HARRISON *J. Clin. Investigation* 9, 693-703(1931), cf. *C. A.* 25, 1283.—Deaths of total solid and K were made on skeletal muscle, cardiac muscle from both ventricles and on liver and kidney in subjects dying without cardiac disease, with congestive heart failure without K and with congestive heart failure when K was administered during life. The controls showed a higher K content in both wet and dry tissues than the heart failure cases. The administration of K to the heart cases increased the tissue K, but the difference was most striking with skeletal muscle and least in heart muscle. J. B. BROWN

The toxicity and balantidicidal action of "di-hydranol" in guinea pigs. NORMAN A. DAVID AND CHAUNCEY D. LEAKE *Proc. Soc. Exptl. Biol. Med.* 28, 196-9(1930).—"Di-hydranol" (2,4 dihydroxyphenyl n heptane) caused death in the guinea pig in about 4 days with symptoms of loss of wt., diarrhea and lethargy at an av. single dose of 400 mg. per kg. of body wt. It cures *Balantidium coli* infested guinea pigs when given in doses of 75-100 mg. per kg. at 1-2 day intervals, until the total amt. administered is 200-225 mg. per kg. of body wt. Doses slightly higher than the curative range may be toxic. C. V. BAILEY

Value of hexuronic acid in the treatment of Grave's disease with suprarenal cortex. DAVID MARINE, EMIL J. BAUMANN AND BRUCE WEBSTER. *Proc. Soc. Exptl. Biol. Med.* 28, 327-9(1931).—Marked clinical improvement followed the oral administration of a glycerol ext. of fresh ox suprarenal cortex in patients with Grave's disease. A hexuronic acid, readily destroyed by exposure to air, had been found in suprarenal cortex (cf. *C. A.* 23, 1423). Large doses of hexuronic acid concentrate, 400 mg. twice daily, given by mouth did not relieve the symptoms of Grave's disease and did not add to the therapeutic value of the suprarenal cortex ext. C. V. BAILEY

Effects of the halogenation of hydroxyquinoline on biological activity. H. H. ANDERSON, N. A. DAVID AND DOROTHY A. KOCH *Proc. Soc. Exptl. Biol. Med.* 28, 484-5(1931).—Toxicity increased with halogenation of hydroxyquinoline and in proportion to the at. wt. of the halogen. The addn. of both I and Cl to hydroxyquinoline caused a considerable increase in toxicity which was slightly reduced by the addn. of a solubilizing group. The balantidicidal action in naturally infested guinea pigs was also increased by halogenation. In monkeys naturally infested with *Endameba histolytica*, 900 mg. per kg. of body wt. of iodochlorohydroxyquinoline given in divided doses over a 6 weeks' period eradicated the amebae. C. V. BAILEY

Blood-sugar response to intravenous insulin in normals and in diabetics. WM. S. COLLEN'S AND HAROLD G. GRAVEL. *Proc. Soc. Exptl. Biol. Med.* 28, 487-9(1931).—Equivalent intravenous doses of insulin on the basis of body wt. produce a much greater depression in the blood sugar level in diabetic than in normal man. The absence of an insulin-inhibiting substance in the blood of diabetics is inferred. C. V. BAILEY

Parathyroid and fluoride hypocalcemia. R. GERSCHMANN *Compt. rend. soc. biol.* 104, 411-2(1930).—See *C. A.* 25, 348. E. J. C.

Parathyroids and hyperglucemic curves. R. GERSCHMANN *Rev. soc. Argentina biol.* 6, 35-9(1930).—See *C. A.* 25, 348. E. J. C.

Bromide poisoning with the picture of typhus abdominalis. HEINRICH MAYRHOFER AND ALFRED FESZLER *Wiener klin. Wochschr.* 43, 1315-6(1930).—A case is described. D. B. DILL

The comparative effects of diuretics containing mercury. ELISABETH BERGER. *Wiener klin. Wochschr.* 43, 1505-8(1930).—The diuretic properties of NH_4Cl , NH_4Br and HCl are due to acid formation in the organism. They are used together with *salyrgan* and *decholin* when these alone are ineffective. D. B. DILL

The effect of age upon the action of thyroxine. EMIL ABDEHARDEN AND ERNST WERTHEIMER *Z. ges. exptl. Med.* 68, 1-19(1929).—Loss of weight and toxicity are more marked in older or heavier guinea pigs and rats when given thyroxine than in younger or lighter animals. The glycogen losses are more marked in the older animals. No definite differences were made out with regard to gaseous and N metabolism at different ages, following administration of thyroxine. Increasing the dosage of thyroxine from 0.3 to 3 mg. per day in young animals results in but very little increase in toxicity. F. L. DUNN

Ketcuria in experimental hyperthyroidism. I. ABELIN AND A. JORDL. *Z. ges. exptl. Med.* 68, 20-31(1929).—Following thyroid ext. feeding or injections of thyroxine in rats there is an excretion of acetone, diacetic acid and β hydroxybutyric acid in the urine, the last being in greatest amt. Maximal excretion occurs on the 10-14th day. It is much less than that observed following phlorhizin or alkaloids. F. L. DUNN

The influence of thyroxine on the blood iodine in myxedema. B. EISLER AND

A. SCHITTENHELM *Z. ges. exp. Med.* 68, 487-92(1929) —The I content of the blood is markedly lowered in myxedema. Thyroxine administration (oral or intravenous) raises the blood I above the normal values for a short period after which it remains normal. The blood I in normal individuals is not altered by thyroxine feeding. F. L. DUNN

Mechanism of action of small doses of salts and "Trunccek-serum." K. I. KOTELNIKOV *Z. ges. exp. Med.* 69, 297-320(1930) —Subcutaneous injection of inorg. salts (5-15-25 mg.) and Trunccek serum results in definite changes in the acid base balance measured by alveolar air CO_2 . F. L. DUNN

Intravenous injection of ethyl alcohol for anesthesia. I. I. NITZESCU *Compt. rend. soc. biol.* 104, 25-8(1930) —See C. A. 24, 5865. G. C.

Action of arsenite on tissue respiration. ALBERT SZENT GYÖRGY *Biochem. J.* 24, 1723-7(1930) —Arsenite strongly inhibits respiration of minced liver tissue, but has no effect on O activation, and little effect on H activation. In yeast, arsenite in higher concns. inhibits respiration but has no effect on alc. fermentation. B. H.

Ethylene and carbon dioxide. JAMES TAYLOR GWATNEY AND M. HILLEL FELDMAN *Dental Cosmos* 72, 1164-6(1930) —The use of C_2H_4 , N_2O , CO_2 , and O_2 for general anesthesia is described. JOSEPH S. HIEBURN

Effects of strontium administration on the histological structure of the teeth of rats. HENRY KLEIN, J. ERNESTINE BECKER AND E. V. MCCOLLUM *J. Dental Research* 10, 733-8(1930) —A ration, adequate in all respects except its Ca content and contg. Sr chemically equiv. to 0.6% Ca, produced a proliferation of dentinoid, i. e., the uncalcified ground substance, when fed to rats. JOSEPH S. HIEBURN

Contraction of the extrinsic muscles of the eye by choline and nicotine. W. S. DUKE-ELDER AND P. M. DUKE-ELDER *Proc. Roy. Soc. (London)*, B107, 332-43(1930) —Choline, acetylcholine and nicotine produce a slow, tonic contraction of the normal extrinsic muscles of the eye, which is unaffected by atropine, increased by adrenaline and abolished by nicotine plus curare. JOSEPH S. HIEBURN

Experimental addiction of animals (monkeys) to opiates. LAWRENCE KOLB *U. S. Pub. Health Repts.* 46, 698-726(1931) —The relative effects in the order of their severity of the 3 drugs given to monkeys in one toxic dose or in smaller daily doses were: toxicity, heroin was decidedly most toxic, morphine least toxic, general health, codeine was decidedly most harmful, heroin least harmful, dependence producing properties, morphine was most potent, codeine very slight or nonexistent. J. A. KENNEDY

Intraperitoneal iron. CLYFFORD G. GAULES AND HENRY WORTH N. SANFORD *Am. J. Diseases Children* 41, 53-61(1931) —Fe injected intraperitoneally into rabbits becomes deposited in the liver rather than in the spleen. The hemoglobin and erythrocyte contents of the blood remain unchanged. The intraperitoneal administration of Fe to children with secondary anemia appears to be a safe and easy way of introducing Fe into the organism and creating a reservoir upon which the body may draw. Following the intraperitoneal injection of 5-8 mg. of Fe, as colloidal $\text{Fe}(\text{OH})_3$, at 3-day intervals, the hemoglobin content of the blood and the no. of erythrocytes become markedly increased. The injections are combined with exposure to ultra-violet light or with blood transfusions. E. R. MAIN

The blood constituents in acute rheumatism before and after salicylate treatment. EVELYN M. HICKMAN AND SIDNEY H. EDGAR *Arch. Disease in Childhood* 5, 387-96(1930) —The urea, non protein N and uric acid contents of the blood are somewhat higher than normal in acute rheumatism. The values tend to remain high during treatment with salicylate. The urea tends to become increased after the treatment while the administration of the same quantities of salicylate to normal children does not affect the urea content of the blood. The treatment with salicylate may further decrease the degree of renal efficiency in acute rheumatism. E. R. MAIN

Quantitative studies on local anesthetics. A. RABBENO *Boll. soc. ital. biol. sper.* 5, 1094-8(1930) —The method used consisted in the application, to the dorsal skin of the leg of the decerebrated frog, of a rectangular piece of cotton saturated with the soln. under examn., this was left for a definite time, and the reflexes were tested with disks of paper wetted with N AcOH. By varying the time of application of the drug, at const. concn., the min. time necessary to obtain the abolition of the reflexes is detd. and with successive tests at varying concns. a curve of the min. time of anesthesia as a function of the concn. of the drug is obtained. The curve is bound by 2 limiting values of concn.: the one which provokes immediate anesthesia (1 min.) and that which acts after several hrs. of application. Pyrrol alkyl or aryl ketones (α acetyl, α propionyl, α butyryl, α benzoylpyrrole). The curve which graphically represents the manner with which the min. time of anesthesia of pyrrol ketones varies with varying concns. of the drugs is

an arc of an hyperbola, represented by the equation $y = a + (b/x)$ (1), where $y = \text{min. time}$, $x = \text{concns.}$, a and b are consts. When $y = 0$, $x = b/a$, that is, for zero time, the curve intersects the axis of x at a concn equal to $-b/a$. For concns equal to or greater than this, anesthesia is almost immediate. From the intersection at x , the curve rises with varying rapidity as indicated by the numerical value of the const a (acetylpyrrole —58.8, propionyl and butyryl —4). For values of y between 60 and 140 min the relation between the concns is a const and the conclusion is that the anesthetic activity of the acetyl-, propionyl-, butyryl and benzoylpyrrole increases as 1 10 110 250. Referred to cocaine, the activity of the acetyl and propionyl derivs. is less, while that of the butyryl and benzoyl derivs. is 8 times greater. Cocaine and cocaine substitutes (cocaine, stovaine, alpine, novocaine, eucaine B and tutocaine HCl). The curves for the min time for these drugs resemble those of the pyrrol ketones. Cocaine gives values corresponding to equation (1), novocaine to $xy = b$, and the others to $y = b/(x - a)$. It was observed that the ratio alpine/cocaine, tutocaine/cocaine, stovaine/cocaine and eucaine/cocaine is const. only for very high values of min time (150–200 min.) while for novocaine/cocaine it is 60–150 min. For lower values of min time (10–30, 60 min.), different numerical values are obtained. Conclusion: The activity of the local anesthetics used, tested by the method described, is less than that of cocaine. For brief intervals of anesthesia, alpine has just slightly less, tutocaine and stovaine has $2/3$ – $3/4$, eucaine $1/2$ and novocaine $1/20$ – $1/80$ the activity of cocaine. PETER MASUCCI

Chloroform narcosis and radon. P. MASCHERA. *Boll soc ital biol sper* 5, 1167–72(1930).—White mice narcotized with CHCl_3 and kept in an atm rich in Ra emanations showed more rapidly phenomena of excitation and the narcosis intervened more quickly than in the control mice that were also narcotized but kept in ordinary air. Also on awakening, the radioactivated animals required a longer time to reestablish normal functions than the controls. PETER MASUCCI

Uroselectan in pyelography. JUAN SALLERAS. *Semana med* (Buenos Aires) 1931, I, 570–1.—Uroselectan was given intravenously. In most cases satisfactory results could not be obtained. A. E. MEYER

Irradiated ergosterol in the treatment of tuberculosis. F. MEERSEMAN AND G. TRICAULT. *Semana med* (Buenos Aires) 1931, I, 595–6.—In 11 human cases, improvement was observed. Tuberculous guinea pigs treated with irradiated ergosterol lived longer than the controls, and the lesions were found to be less severe. A. E. MEYER

The treatment of asthma with sulfarsenol. FÉLIX J. LICEAGA. *Semana med* (Buenos Aires) 1931, I, 636–51.—Clinical cases of successful treatment are reported. A. E. MEYER

Clinical investigations into the effect of intravenous injections of insulin. V. The action of adrenaline on the stomach secretion. KAJ ROHOLM. *Acta med. scand.* 74, 359–69(1931), cf. C. A. 25, 353.—It was found that injections of insulin stimulate gastric secretion and that this stimulation was really an adrenaline effect. However, adrenaline injections have no effect upon the secretion of the fasting stomach. S. MORGULIS

The effect of colloids on histamine action. BEHREND BEHRENS. *Biochem. Z.* 231, 92–4(1931).—Small amts. of kaolin or talcum do not affect the action of histamine but large amts. may cause inhibition. S. MORGULIS

The fate of parenterally administered sulfur. R. MEYER-BISCH AND F. TECHNER. *Biochem. Z.* 231, 110–2(1931).—No changes either in the water content or in the hydrolyzable and total S of the liver have been found in rabbits receiving S parenterally. The changes in S metabolism must therefore be attributed to processes occurring in more peripheral organs, such as cartilage, skin, etc. S. MORGULIS

Influence of insulin on blood and organ lipids. LAVINE H. PAGE, L. PASTERNAK AND MARIE L. BURT. *Biochem. Z.* 231, 113–22(1931).—The most pronounced effect of insulin is manifested in the phosphate content of serum, which diminishes 30%. The cholesterol contents of the blood, kidney and brain are higher than those in the non-insulinized controls. The aliphatic acid contents of brain, blood and serum seem to diminish under insulin, but the I_2 nos. of the acids increase in some (brain, liver) and decreases in other tissues (blood, serum). S. MORGULIS

The relative toxicities of some organic salts of triethyl lead hydroxide. HENRY GILMAN AND O. M. GRUBIT. *J. Pharmacol.* 41, 1–4(1931).—The m. l. d. intravenously to rats, and the m. l. d. for the first 3 intramuscularly to rabbits are given for the following: triethyl lead salicylate, triethyl lead phenyl acetate, triethyl lead furoate, triethyl lead *p*-toluenesulfonate, triethyl lead *p*-aminobenzoate, triethyl lead furyl-acrylate. All the compds. were extremely irritating and produced nervous symptoms. C. RIEGEL

Notes on the importance of exact preparation of tincture of digitalis and on the number of pigeons in the pigeon-emesis method. P. J. HANZLIK, A. B. STOCKTON AND S. S. DAVIS. *J. Pharmacol.* 41, 5-10(1931).—The min emetic dose (m em d) of digitalis tincture varied with the length of time allowed for extn of the digitalis leaves. The practicability of using as the m em d that amt. causing emesis in 2 out of 3 pigeons is confirmed by detns. on larger groups. C. RIEGEL

Some clinical actions and therapeutic uses of racemic synephrine. A. B. STOCKTON, P. T. PACE AND M. L. TAINTER. *J. Pharmacol.* 41, 11-20(1931).—Racemic synephrine raised the blood pressure in patients when 0.2 g intramuscularly or 0.05 g intravenously was given the rise being from 25 to 30 mm Hg, and lasting 1/4 to 1 hr. It was ineffective in asthma, and did not prolong the action of procaine in local anesthesia. Applied to the nasal mucous membrane it produced shrinkage of the mucosa without irritation, engorgement or swelling. Salts of synephrine are stable and could be sterilized by heat. C. RIEGEL

The effect of mechanical constriction of the hepatic veins upon the anticoagulant action of Witte's peptone. ELIZABETH CRANSTON AND O. R. CALLET. *J. Pharmacol.* 41, 65-70(1931).—Constriction of the hepatic veins previous to injection of the peptone produced a delay in the appearance and a reduction in the intensity of its anticoagulant action, and a less marked reduction in platelet count. C. RIEGEL

Apomorphine tolerance and its relationship to morphine tolerance. FRANK CO TUI. *J. Pharmacol.* 41, 71-82(1931).—Dogs habituated to apomorphine develop a tolerance to its emetic action, but not to its other effects. Dogs tolerant to apomorphine are tolerant to the emetic action of morphine, and vice versa. There is no cross tolerance as regards other effects. Neither morphine- nor apomorphine-tolerant animals are tolerant to the emetic action of pilocarpine. C. RIEGEL

Glycogen storage in the white rat when fed the roots of *Arctium leppa*. JOHN C. KRANTZ, JR., AND C. JELLEFF CARR. *J. Pharmacol.* 41, 83-7(1931).—Glycogen storage in the liver was increased in rats fed the roots or the carbohydrate prep. from them, over that in rats fed only on a basal diet of cacao butter. C. RIEGEL

The influence of insulin on the formed blood elements, on the sedimentation rate of the erythrocytes, and on the bleeding and coagulation time. GIULIO LEGA. *Arch. farmacol. sper.* 51, 1-23(1930).—In the blood of the diabetic there exists in general a diminution of hemoglobin and red cells, a slight and inconst. increase in leucocytes and bleeding time, and an almost const. increase in coagulation time and sedimentation rate. The graver the condition of the diabetic is, the more marked are these changes, though they are not directly related to the degree of glucemia. Dietetic treatment resulting in an improvement of the general condition of the patient causes an increase in erythrocytes and hemoglobin, but it does not influence the leucocyte formula, coagulation time, bleeding time and sedimentation rate. Insulin treatment resulting in a notable improvement of the general condition of the patient has a remarkable influence on the factors under examn., since coagulation time, bleeding time, sedimentation rate and no of formed blood elements returned rapidly and almost constantly toward physiol. conditions. G. SCHWOCH

The behavior of experimental glucemia in cases of diabetes mellitus treated with insulin. GIULIO LEGA. *Arch. farmacol. sper.* 51, 33-40(1930).—When 20 g. of glucose in 200 cc. water was given to healthy persons, the blood sugar, which was detd. every 30 min for 3 hrs., rose to a max. during the 1st hr. after ingestion of the glucose. Then it decreased and finally reached values below the normal level. When the same test was carried out in diabetics, the rise in blood sugar was much greater, the values remained at an elevated level for a considerably longer period, the decrease was much slower, and the original level was not reached at the end of the 3rd hr. The grave cases presented the strongest deviations from the normal curve. The more a case improved under insulin treatment, the more closely the glucemic curve approached the normal type. The improvement in the condition of the patient is due to the improved functioning of the pancreas brought about under the influence of insulin. Thus the test described may serve as a method of detg. the functional ability of the pancreas with respect to its internal secretion. G. SCHWOCH

Is codeine without influence on the blood pressure? LUIGI SCREMIN. *Arch. farmacol. sper.* 51, 97-110(1930); cf. C. A. 24, 4643.—Intravenous injections of small doses of codeine caused an increase in blood pressure in rabbits and guinea pigs. In the cat the blood pressure was raised by very small doses of codeine, somewhat larger doses, which, however, were still far below the min. convulsive dose, caused a considerable decrease in blood pressure. G. SCHWOCH

The action of photodynamic substances on the carbohydrate metabolism. II.

GABRIELE MONASTIRIO *Arch. farmacol. sper.* 51, 111-21 (1930).—In an earlier study M had investigated the influence of tryptaflavine on the blood sugar of rabbits. In this work its effect on the blood sugar of dogs was studied. Intravenous doses of 5 mg per kg were without effect on the blood sugar. Ten mg caused a pronounced hyperglucemia, while 40 mg caused hypoglucemia after about 7 hrs. Tryptaflavine is quite toxic for dogs, and therefore the 40 mg dose was given in single doses of 10 mg each at intervals of 1 hr. This method of subdividing the dose is probably the reason why the hypoglucemia was preceded by a hyperglucemia. The symptoms of hypoglucemia consisting of generalized tremors and sometimes convulsions disappeared after glucose injection, but the animal remained in poor condition because of the intoxication caused by the tryptaflavine. When rabbits injected with 20 mg tryptaflavine per kg were irradiated by ultra violet rays the hyperglucemia produced in them was much more intense than when they were kept in the dark after injection of the same dose. Other expts were concerned with the influence of antiphotodynamic substances injected simultaneously with or after, the administration of tryptaflavine. When small doses of resorcinol totalling about 0.5 g were injected into dogs, which a few hrs previously had received sufficiently large doses of tryptaflavine, the hyperglucemia developed normally, but the onset of hypoglucemia was greatly retarded. When the resorcinol was injected at the beginning of the expt., the onset of hyperglucemia was retarded. Eosin injected subcutaneously or intravenously did not alter the blood sugar of normal rabbits. The hyperglucemia induced in rabbits by adrenaline or subcutaneous injection of glucose was intensified after eosin administration. On the other hand, the hypoglucemia producing effect of insulin was attenuated by eosin. Of some other acridine derivs studied, acridine orange and, to a slight extent, flavic acid are capable of producing hyperglucemia. Rivanol was without effect. Inconst. results were obtained with acridine yellow and acridine red. G. SCHWACH

Changes in the osmotic pressure of the organs after intravenous injection of urea. LUIGI SCERBIN *Arch. farmacol. sper.* 51, 121-14 (1930), cf. *Messina C. A.* 24, 657.—The expts were carried out on 8 rabbits to which quantities of urea varying from 2.10 to 0.80 g per kg were slowly administered by intravenous injection of a 13% soln. The animals were bled to death 90 sec after the injection was completed and the osmotic pressure of the blood serum and various organs was detd by the cryoscopic method. In normal rabbits Δ rises in the organs examd in the following order: serum, brain, muscle, heart, kidney, liver. In the exptl animals Δ generally increased with increasing doses of urea. For a certain dose, serum always showed the lowest values, slightly higher values were found in the brain, while muscle and liver gave the highest figures. However, the relative increase was greatest in the serum and was smallest in the kidney and liver. The results obtained are discussed in connection with the pathogenesis of uremia, but no definite conclusions are drawn. G. SCHWACH

Influence of some carbohydrates on the toxicity of potassium cyanide. P. MORFITT AND G. MASCOLINO *Arch. farmacol. sper.* 51, 135-40 (1930).—Mixts of 4% KCN soln and glucose soln (10-30%) were allowed to stand at 37° for 24 hrs. When such mixts. in quantities corresponding to 2, 4, and 10 times the m. l. d. of KCN were injected intramuscularly into pigeons, the animals remained alive and well. It is calcd from the quantities of glucose added that 1 g KCN is neutralized *in vitro* by 15 g glucose or less. After a certain period of time the KCN-glucose solns assume a distinctly ammoniacal odor; however, the reactions of Wortmann and Pagenstecher-Schönbein remain positive. In attempting to hydrolyze the glucose in the glucose-KCN complex a mixt. of KCN and glucose soln was treated with yeast. After the treatment the mixt. does not reduce Fehling soln., which shows that KCN, after combining with glucose, has lost its antienzyme properties, however, its toxicity was not restored. Similar expts. in which glucose was replaced by other carbohydrates showed that fructose, galactose, mannose and lactose also neutralized KCN, while dextrose, sucrose and gum arabic did not. Starch was only very slightly effective as a neutralizing agent. G. SCHWACH

The action of yohimbine on the blood pressure and the reversal of the blood-pressure action of adrenaline and adrenalone by yohimbine. MASASHI YAMAUCHI, *Okayama Igakkaï-Zasshi* 40, 1220, *Chem. Zentr.* 1929, II, 3029.—Small doses of yohimbine cause an increase in blood pressure by exciting the vascular muscle itself. The blood-pressure reducing effect of larger doses of yohimbine seems to be due to the ability of the poison to render the sympathetic fibers insensitive for the adrenaline, which always circulates in the organism, therefore only the action of adrenaline on the vasodilators becomes manifest, which then surpasses the exciting action of yohimbine on the vascular muscle. By a previous treatment with yohimbine the action of adrenalone is inhibited more strongly than that of adrenaline. G. SCHWACH

Reversible coagulation in living tissue. I WILDER D BANCROFT AND J. E. RUTZLER, JR. *Proc Natl Acad Sci* 17, 105-11(1931)—Anesthesia and anaphylactic shock are due to the reversible coagulation of protein in certain nerve colloids. Intravenous injection of a 10% NaCNS soln., which peptizes protein, hastens the return of consciousness of rabbits under ether, amylal or morphine anesthesia and prevents death from strychnine or histamine. It prevented anaphylactic shock in rabbits sensitized by subcutaneous injection of egg white sol. KCNS cannot be used because of its high toxicity. D S SEARLE

The influence of insulin and adrenaline on the sugar content of the spinal fluid. J LUKÁCS. *Magyar Orvosi Archivum* 30, 591-7(1930)—The sugar of the spinal fluid is decreased after the injection of insulin. Decrease in blood sugar sets in sooner than that in spinal fluid. Preceding the decrease in the sugar of the spinal fluid a temporary increase takes place, this is not true for blood sugar. Adrenaline raises the sugar in the spinal fluid and also in the blood, however, the increase in blood sugar takes place first. H TAUBER

The action of histamine on human blood vessels. J POGÁNY. *Magyar Orvosi Archivum* 32, 51-6(1931)—0.001-0.002 mg of histamine intravenously injected into the human vein causes a rise in venous pressure, which sets in immediately after injection but only in veins affected first by adrenaline. 0.01-0.02 mg histamine also raises venous pressure. The rise sets in after 25-60 sec. and is manifested by accompanied warmth. This can be eliminated if the vein is shut off in the capillary region. H TAUBER

The action of histamine on human adrenaline sensibility. J POGÁNY AND G PILAU. *Magyar Orvosi Archivum* 32, 57-61(1931)—The adrenaline sensibility is increased considerably 30-50 min. after subcutaneous injection of small doses (0.005-0.01 mg per body wt.) of histamine. This effect is attributed partly to the dilatation of peripheral small vessels and partly to the characteristic histamine distribution in the blood. H TAUBER

Induction of lymphocytosis and lymphatic hyperplasia by means of parenterally administered protein. BRUCE K WISEMAN. *J Exptl Med* 53, 499-510(1931)—In this study 2 animals received chick "embryonic ext." intraperitoneally, 2 intravenously and 1 subcutaneously, 2 received egg albumin intravenously and 1 subcutaneously, 1 received normal horse serum intravenously and 1 received normal salt intravenously, each injection carried approx. 50 mg of protein in 60 cc normal salt soln. The injections caused an increase of lymphocytes in the peripheral blood varying from 23 to 139%. It seems probable that the degree of response is conditioned upon the type of protein used. At autopsy the lymph nodes and spleen showed hyperplastic changes. The thymus did not participate in the hyperplasia. C. J. WEST

I—ZOOLOGY

R. A. GORTNER

The oxidation-reduction potential in slugs (*Agriolimax agrestis*). EUGÈNE AUBEL AND ROBERT LÉVY. *Compt rend soc biol* 105, 358-9(1930)—Oxidation reduction indicators were injected into living slugs. The oxidation reduction potential figures for slugs were identical with those formerly obtained for caterpillars. The physiol. r_H in aerobiosis is about 20, in anaerobiosis, it is lowered to a limit between 5 and 6. B. C. BRUNSTETTER

The action of placenta on the metamorphosis and on the development of the smooth muscles in tadpoles. G. GANFINI. *Boll soc ital biol sper* 5, 949-53(1930)—The larvae of *Bufo vulgaris* were fed with raw beef, raw and cooked human placenta, aq ext of placenta and planktons. Raw cooked or exts of human placenta contained substances which accelerated the development and metamorphosis of tadpoles and activated the development of the smooth muscles. The amt. of these substances was, however, greatest in the beef. PETER MASUCCI

The importance of arginine-phosphoric acid in the metabolism of the reacting crustacean muscle. EINAR LUNDSGAARD. *Biochem Z* 230, 10-8(1931)—The metabolism of the reacting crustacean muscle is altered exactly like that of the frog muscle when its lactic acid production is stopped by poisoning with CH_3I COOH, whereby the arginine- H_2PO_4 of the crustacean muscle behaves just as the creatine- H_2PO_4 of the frog muscle. It is therefore concluded that the hydrolysis of either arginine- or of creatine H_2PO_4 must be considered as a source of energy. The crustacean muscle is very much less sensitive to CH_3I CO $_2\text{H}$ than is frog muscle. S. MORGULIS

Hydrogen-ion concentration in a unicellular body. S. DATE. *Compt rend. soc.*

biol. 106, 80-83(1931) —The pH of the protoplasm of *Amoeba proteus* in the culture ($pH = 8.0$) is 6.4 to 7.3 (it is more acid during movement than during repose), the pH of the nucleus is a little higher than that of the protoplasm. When the exterior pH varies from 5 to 9, it does not greatly influence the interior pH . ALBERT L. RAWLINS

A radioactive indicator method for estimating the solubility of acid lead arsenate within the alimentary tract of the silkworm. F. C. CAMPBELL AND CHARLES LUKFMA. *J. Econ. Entomol.* 24, 88-94(1931) —This study describes the use of a radioactive indicator method for the detn. of the amt. of Pb evacuated by the silk worm (*Bombyx mori*) after being fed known dosages of PbH_2AsO_4 and $Pb_3(PbOH)(AsO_4)_2 \cdot 11H_2O$. An arsenate of Th B, a radioactive isotope of Pb, is mixed with known quantities of Pb arsenate. The mixt. is then led to an insect. The Th B accompanies the Pb wherever it goes in the body and the ratio of Th B to Pb will remain the same as in the original mixt. regardless of the reactions into which the arsenate may enter. The prepn. which was fed to the insect was tested in an electroscope to obtain the quantity of Pb present. The ash of the excrement was similarly tested and from this information the quantity of Pb, and therefore of the Pb arsenate evacuated, was obtained. Results: At least 25% of a moderately lethal dose of PbH_2AsO_4 (0.12 mg. per g. body wt.) passes into soln. within the intestine during the survival period. Basic lead arsenate was much less sol. within the intestine. C. H. RICHARDSON

The catalase content of the Colorado potato beetle during metamorphosis. DAVID F. FINK. *J. Agr. Research* 41, 691-6(1930) —The variation in the activity of catalase during the metamorphosis of the Colorado potato beetle, beginning with the mature larva, was ascertained by making daily detns. of the quantity of O_2 per g. of tissue evolved from H_2O_2 by the action of catalase during prepupal and pupal development. Analysis of the resulting data indicated a reduction in catalase of 10% on the first day of histolysis and 10% on the second, followed on the third day of histolysis by an increase which, when the pupa was formed, reached the max. of 6% above that of the mature larva. The decline in content of catalase during pupal development was continuous thereafter until a reduction of 51% below that of the mature larva was reached with the formation of the adult. The curves obtained for catalase content and those previously reported for O_2 consumption and excretion of CO_2 (C. A. 19, 2536) were similar for the prepupal period of histolysis. During pupal development, however, the catalase curve declined rapidly whereas the curves for O_2 and CO_2 followed the usual growth curve. W. H. ROSS

Phosphorus distribution, sugar and hemoglobin in the blood of fish, eels and turtles. C. M. MCCAY. *J. Biol. Chem.* 90, 497-505(1931) —The distribution of P in blood fractionated by Youngburg's method (C. A. 25, 1857) was detd. by a modification of the method of Kuttner and Cohen (C. A. 22, 806); $HClO_4$ was used in place of H_2O_2 for the oxidation. The total P in 100 cc. whole blood of pike is 87 mg., carp 92 mg., bullhead 77 mg., *Chelydra serpentina* 49 mg., compared with beef blood 22.0 mg. Figures for the P distribution, blood sugar and Fe are given. During spawning the hemoglobin and reducing sugar in the blood of lamprey eels decreases while the total P is unchanged. K. V. THIMANN

Physicochemical properties of crocodile blood. D. B. DILL AND I. T. EDWARDS. *J. Biol. Chem.* 90, 515-30(1931) —The concn. of protein in the serum and of hemoglobin in the blood of crocodiles is lower than in man. The buffer value of the blood is also much less, although that of the serum protein itself is higher than in man. The change in strength of acidity on oxygenating crocodile hemoglobin is twice as great as in human blood, and correspondingly the effect of pH change on the O_2 affinity is twice as great. The content and distribution of water, Na, K, Ca, Mg, Cl, HCO_3 , phosphates, lactate and proteinate in the plasma and cells were also detd. K. V. THIMANN

Effect of inorganic salts on photic orientation in *Alloobophora foetida* (Sav.). VI. Magnesium sulfate, nitrate and chloride. IKUTARO NOMURA AND SHINRYO OHSUCHI. *Science Repts. Tôhoku Imp. Univ.* [4], 5, 669-89(1930) —A large no. of data are given to show that immersion of the worm in solns. of $MgSO_4$, $MgCl_2$ and $Mg(NO_3)_2$ alters the degree of orientation to light. The forward crawling mechanism in the brain and ventral nerve cord is also affected. The effects are complex. K. V. THIMANN

Effect of light on porphyrin from the integument of the earthworm *Alloobophora foetida* (Sav.). SATARO KOBAYASHI. *Science Repts. Tôhoku Imp. Univ.* [4], 5, 763-78(1930) —The worms were extd. with 20% HCl alc., centrifuged, the filtrate was neutralized and extd. with H_2O . The well-washed ext. was treated first with glacial HIO_4 , then with 20% HCl, re-extd. with H_2O , each time, washed and evapd., the whole operation being carried out in the dark. A yield of 1.5 g. was obtained from 2 kg. worms. The spectrum of the porphyrin in 20% HCl alc. slowly fades on illumination, but in 25%

HCl, in alc. AmOH , PbO or 6% HOAc alc., the 4 bands move slightly toward the violet and 2 new bands appear in the red at 1649 and 670 μ . With the appearance of the new bands the intensities of the others decrease. The spectrum agrees approx. with that given by Fischer and Schaumann (*C A* 17, 3349), and is entirely different from that of hematoporphyrin. The changes occur very rapidly and support the view that these pigments of the erythrocytes are photodynamically sensitive. K. V. THIMANN

Analysis of the proteolytic enzymes of *Maja squinado* by the adsorption method. J. J. MANSOUR BEE. *Proc. Acad. Sci. Amsterdam* 33, 858-70 (1930).—The ext. of the small intestine contains a protease, carboxypolypeptidase, aminopolypeptidase and dipeptidase. The pH optimum of the protease of the raw ext. is 6.0 on casein and 6.1 on gelatin, while that of the purified protease is 7.4 and 8.1. The raw ext. is not activated by HCN, H_2S or enterokinase, while the purified ext. is activated by the latter. Separation of the enzymes was achieved by differential adsorption on kaolin and on fuller's earth. The protease was purified by repeated adsorption on Fe(OH) . J. J. WILLAMAN

12—FOODS

P. C. BLANK AND H. A. LEPPER

Chemical and physical methods in the examination of food products. EGGERT. *Pharm. Presse* (Russ. prakt. Heft) 1931, 20-1.—An address is given which deals with some of the more recent methods. W. O. E.

International agreement for the unification of the presentation of analytical results on foodstuffs for man and animals. ANON. *Ann. Jals* 24, 69-75 (1931).—Final text of the International agreement for the unification of the presentation of analytical results established as a result of the International conference held in Paris on May 13, 1929, by which a unanimous decision was reached. A. PAPINEAU COUTURE

Determination of water in foods, condiments, etc. B. BLEYER AND W. BRAUN. *Z. anal. Chem.* 83, 241-68 (1931).—Water may be detd. (a) by thermic expulsion and weighing the residue, (b) by distn. with a high boiling liquid which is immiscible with water and measuring the vol. of water in the distillate, (c) by treatment with CaCl_2 and measuring the vol. of C_2H_2 formed and (d) by detn. of the diast. const. These methods are discussed, and a no. of typical food ingredients and food products were examd. comparatively. Conclusion: There is no one method which is always preferable, but preliminary tests will show which particular method is best for a given material. W. T. H.

Biochemical data on wheat harvested at the experiment station of Bordenave. CARLOS ALBIZZATI. *Rev. facultad cienc. quim. (Univ. La Plata)* 7, Pt. 2, 53-7 (1930).—Analysis of wheat in the processes of grinding and baking led to the development of an index designated as the "value of utilization." Any wheat having a utilization value of 95 or above is considered suitable. Before discarding a wheat variety because of its low utilization value agronomical and ecological expts. should be made with a view to a possible increase in the utilization value. No conclusions are drawn with regard to bread making qualities and the diastatic properties of the wheat because of the insufficient no. of expts. performed. B. S. LEVINE

Artificial refrigeration in mill industries and its influence on the quality of flours. S. GOMORY. *Magyar Mennak Epiteszseglet Kozlonye* 64, 387-92 (1930).—The heat of the rollers has a bad influence on the chem. and phys. properties of flours. Flours ground on cooled rollers contained about 0.5% more moisture and had their water-absorbing capacities increased by 1-2 cc. for 100 g. of flour. Bread baked from such flours had a somewhat greater vol. and was more homogeneous. The glutenin content of flour did not change, but its quality changed because of its increased moisture content. Intensive refrigeration during rolling prevents evapn. of moisture from wheat, formation of pulpy knots in the flour and eventual fermentation. The flour yield of wheat ground under cooled rollers increased, since peel particles are more easily sepd. from starch and aleuron particles when the moisture content is higher. S. S. DE FINAY

Analytical control of imported flours. JOS. KULMAN. *Chem. Listy* 25, 53-9 (1931).—The methods for analyzing flours suspected of chem. treatment are described. Sp. methods for N oxides, Cl, peroxides, bromates and iodates are given. F. M.

The significance of the acidity in baking. E. ELJON AND L. ELJON. *Chem. Weekblad* 28, 103-5 (1931).—Twenty five g. wheat flour extd. with 500 cc. water contg. 50 mg. tartaric acid gave an ext. (I) of pH 5.9 and with a diastatic activity more than twice as high as the diastatic activity of an ext. of the same flour made with water alone (pH 6.6). Adjusting the pH of I with alkali to 6.6 did not change the diastatic activity of this ext.

By extn. in the presence of tartaric acid the diastase is apparently either extd. in larger quantity or changes to a more active form. F. DE LEEUW

New method for determining the swelling capacity of bread. L. KARÁCSONYI *Magyar Chem Folyóirat* 36, 59-61(1930)—Detn with a viscometer is made as follows: Cut up and moisten 10 g. bread crumbs. After 5 min. press through a sieve 10 cm. in diam. covered with Dufour's silk. Dil. the pulp with water to 100 g. with an exactness of 0.01 g. Pour the homogeneous pulp into an Ostwald viscometer (Lüers and Ostwald, *C. A.* 14, 1167) in a 20° thermostat, mix with an air current and measure the time required for flowing within 0.2 sec. Calc. the av. of 4-5 measurements. The results were well reproducible in the case of wheat breads. Detns. can be made in 30-40 min. as compared with the 24 hrs. of the sedimentation method according to Katz. S. S. DE FINALLY

The distribution of electrolytes in milk and in the dialyzate. ALESSANDRO ROSSI *Boll. soc. ital. biol. sper.* 5, 1060-3(1930), cf. *C. A.* 24, 5359—Milk, 20 cc. samples in celloidion sacs, was dialyzed against 25 cc. of distd. H₂O. Whea osmotic equil. was reached, samples from the inside and outside of the sacs were analyzed for Ca, K and P. The results are given in tabular form. Conclusion: the K in the bound form is rather labile and is dependent on the concn. of the diffusible fraction. PETER MASUCCI

Dialysis of milk. I. MICHIO SAITO *J. Agr. Chem. Soc. Japan* 6, 1087-92 (1930)—Milk was dialyzed for 5 hrs. The outer soln. was evapd. under diminished pressure. The ppt. thus obtained was a mixt. of phosphates and Ca salts. In the filtrate lactose and lactic acid were found. A substance giving no biuret reaction but giving the ninhydrin reaction was found. Boiled milk was also dialyzed. The dialyzable N substances increased markedly, owing to the production of NH₃ by the heating. Y. KIHARA

Does filtered milk keep longer? K. VAS. *Kísérlet. Közlemények* 32, 342-55 (1929), *Chimie & industrie* 25, 453(1931)—Filtration removes the coarser impurities but facilitates the dispersion of bacteria throughout the mass of the liquid, favoring more rapid decompn. of the milk. Of ten samples, in 1 the result was clearly unfavorable, in 3 no difference was noted, in the remaining 6 there was a retardation (up to 11 hrs.) in the rate of coagulation. It would appear that filtration is efficacious only if it is carried out immediately after milking, before the bacteria originally adherent to the larger particles of fecal matter and dirt have become distributed through the milk. A. PAPIEVAU-COUTURE

The distribution of phosphorus in cow milk and a scheme for the separation of phosphatides. RINJIRO SASAKI *Bull. Agr. Chem. Soc. Japan* 6, 52-61(1930)—Milk powder, dried by the Buřlovak drum drier below 70°, was used. The greater part of the phosphatides was more easily sol. in warm alc. or warm acetone than in ether. The existence of much P in the phosphatide fraction has been proved. Y. KIHARA

Commercial sour cream. F. J. DOAN. *Creamery and Milk Plant Monthly* 20, No. 2, 21-2(1931)—A method of mfg. com. sour cream is given. Before the starter is added, the cream is pasteurized for 30 min. at 160°F. Pasteurization at the high temp. yields a cream which will increase in acidity more rapidly during ripening and which will reach max. viscosity in a shorter time. Homogenization of the cream after pasteurizing is desirable, as the sour cream produced from homogenized cream will be thicker in body and smoother in texture than sour cream produced from unhomogenized cream. Increased body of the sour cream may also be obtained by using cream of higher fat content, by adding solids in the form of plain condensed milk, by the use of gelatin and by the use of rennet. The cream is caused to sour by the addn. of 1-2% of lactic starter. The souring is allowed to proceed at 70°F. until 0.6 to 0.65% acid (as lactic) has developed, whereupon the cream is cooled and bottled. A. H. JOHNSON

The basic viscosity of ice cream mix. J. C. HENING *Ice Cream Trade J.* 27, No. 2, 41-5. *J. Dairy Sci.* 14, 84-92(1931)—Mech. agitation of ice cream mixes reduces their viscosity by a partial splitting of fat clumps. The viscosity of mechanically agitated ice cream mixes may be further reduced by passing them through the homogenizer at low pressures. Different conditions of agitation may vary in their effectiveness in splitting fat clumps and introduce a variable factor in the magnitude of the basic viscosity. The term basic viscosity, as now used for ice cream mixes, denotes a value secured under specified conditions and not a correct max. value from the viewpoint of its being the lowest possible viscosity in the absence of fat clusters. A. H. JOHNSON

Determination of the egg content of egg pastes. O. SZAKACS. *Kísérlet. Közlemények* 32, 459-60(1929), *Chimie & industrie* 25, 309(1931)—Digest 2 g. of paste with 100 cc. distd. H₂O for 1 hr. shaking at 15 min. intervals; let stand 30 min.; to 70 cc. of clear soln. add 5 cc. of 5% β -naphthosulfonic acid, coagulate on a water bath at 90°, cool

without shaking and centrifuge. The egg content is obtained by comparing the coagulum obtained with that of a blank detn. As the results are very const., a graduated tube may be used instead of comparison with a blank. A P.-C.

Hawaiian honey. JAR. SVONODA. *Chem. Listy* 24, 462-3(1930) —Hawaiian honey imported into Czechoslovakia during 1927 had a powerful aroma and a saline taste. A complete analysis is given. The salty taste was due to a high ash content (0.47-0.88%), in which the NaCl (0.03-0.09%) and KCl (0.19-0.29%) predominated. FRANK MARSH

Note on the manufacture and analysis of chocolates. RAOUL LECOQ. *Ann. fals.* 24, 11-22, 96-104(1931). See C. A. 25, 153, 1295. A PAPINEAU COUTURE

Refrigeration an essential in candy manufacture. TRESFEE CLARKE. *Refrigerating Eng.* 20, 219-22(1930) —C. outlines the manner in which temp. control has been applied to the candy industry. The temp. of the chocolate mix det. the viscosity of the mix and other properties assoc. with viscosity. The proper rate of cooling of chocolate candy makes it possible to prevent "graying" or dulling. Also the rate of cooling of sugar candies enables one to control the type of crystal, and to obtain the type of crystal desired. Dry rooms are maintained by removing excess moisture from the atm. by cold water or brine circulating in pipes. Maintenance of proper temp. and humidity makes it possible to store candy for 6 months to 1 year without deterioration. A. H. JOHNSON

The disappearance of ozone in cold-storage rooms. ARTHUR W. DWELL. *Refrigerating Eng.* 20, 358-60(1930) —Under practical conditions the rate of decompn. of O_3 follows the equation for monomolecular reactions. With an initial concn. of 1 part of O_3 per million by vol. the absorption in one half hr. by the several food products studied was eggs 0.03 p. p. m., fresh apples 0.33 p. p. m., apples previously exposed to O_3 0.06 p. p. m., fresh meat 0.38 p. p. m., meat previously exposed to O_3 0.20 p. p. m. and cauliflower 0.21 p. p. m. A. H. JOHNSON

A new method for the prevention of mold on marmalades, jellies, fruit pastes, etc. RUDOLF RIPA. *Konserven Ind.* 1931, 171-3 —After jellyfication has begun, but before complete cooling, the surface of the product is coated with a thin layer of pectin soln. The pectin soln. dries to a glaze which prevents mold growth. After complete cooling the container is sealed. FRANCIS P. GRIFFITHS

Blue and green molds of oranges. G. B. TINDALE AND S. FISH. *J. Dept. Agr. Victoria* 29, 101-4(1931) —The atm. of the packing shed and the dust on the grader are the important sources of infection of oranges with blue and green molds. Both molds may spread to sound fruit by contact. The optimum temp. for infection with both types of mold is 70-78°F. With shallow inoculations development of the molds is completely controlled by keeping the fruit at 94°F for 5 days. Oranges, however, will not stand a temp. of 94°F for more than 4 days without developing a bitter flavor. Storage of the fruit at 94°F for 3 days, which gives about 90% control of the molds, is recommended. K. D. JACOB

Organic solvents for aiding the removal of spray residue from waxy or oil-covered fruit. R. L. ROBINSON. *J. Econ. Entomol.* 24, 119-25(1931) —The removal of spray residues from apples which have been sprayed with petroleum oils, or which are covered with an excessive wax formation, is difficult. The addn. of certain org. solvents (alc., $CaH_2(CH_3)_2CO$, kerosene) to the HCl washing soln. renders removal of the residue most effective. Kerosene appears to be the most practicable for this purpose. C. H. R.

A simple test for sulfur dioxide content of dried apricots. W. R. JEWELL. *J. Dept. Agr. Victoria* 29, 90-1(1931) —Add 500 cc. of water to 2 oz. of the minced sample contained in a stoppered bottle and shake thoroughly every 10 min. for 2 hrs. Add a teaspoonful of starch soln., prepd. by boiling a teaspoonful of starch with 2 cups of water, and 18 cc. of 0.1 N I soln., and shake vigorously for 15 sec. If the dark blue color persists at the end of this period, the sample contains less than 14 grains of SO_2 per lb. of fruit. A gray or apricot color indicates the presence of larger amts. of SO_2 . If a green color develops, the test should be repeated with the addn. of a larger amt. of starch soln. The method is designed for rough testing at the packing shed. K. D. JACOB

The supposed relation between the pH value and acid taste of aqueous solutions, particularly wines. P. CRIST. *Ann. chim. applicata* 20, 566-83(1930) —Although there is undoubtedly a relation between the pH value of a wine and its acid taste, other factors influence this taste markedly, such as sugar, EtOH, glycerol, salts, tannins, etc. The undissoc. org. acid radicals also affect this taste, as the addn. of org. acids affects the acid taste noticeably, while it only affects the pH value slightly. Dln. of sour wines with water does not affect the pH value much, but it greatly changes the taste. A. W. CONTIERI

Refrigeration and the fishing industry (FRAN) 13. Electricity in modern dairy plants (DRUX, BRUNNER) 4. Isolation of phytosterolin from wheat embryo (NAKAMURA, ICHIBA) 10. Cherries cultivated in the Modena Province (D'IPPOLITO, ALBERTI) 11D. Biochemistry of tomato pigment (VON EULER, et al) 11D. Water requirements [in creameries] (KEHREN, STOHMEL) 14. Activating metals or alloys [for containers for milk, butter or cheese] (Austrian pat. 121,243) 3. Cascading apparatus for drying cereals, etc (U S pat 1,796,324) 1.

CRUESA, W V. Commercial Fruit and Vegetable Products. New York: McGraw-Hill Book Co., Inc 515 pp \$4.60

JORDAN, EDWIN O. Food Poisoning and Food-borne Infection. Chicago: Univ of Chicago Press 286 pp \$2.50

KNAPP, A W. The Cocoa and Chocolate Industry. 2nd ed. London: I. Pitman and Sons, Ltd 212 pp 7s 6d. Reviewed in *Chimie & Industrie* 25, 534 (1931).

TARUGI, N. Trattato di chimica bromatologica. Milan: F. Vallardi. 460 pp L 50

WOODMAN, ALPHEUS G. Food Analysis. 3rd ed. New York: McGraw-Hill Book Co., Inc 657 pp \$3.60

Treating meal. ERICH STAUDT. Ger 515,843, May 22, 1926. Foodstuffs, especially meal, is sterilized, matured and bleached by treatment with compds contg. ClO_2 . In the example, wheat or rye flour is treated with gaseous ClO_2 mixed with air and CO_2 . The gas is absorbed by the meal.

Improvement of the baking capacity of cereal flours. M. VUK and S. GOMORY. Hung 101,845, Nov 12, 1929. A paste-like pulp is formed of the flour with water. This pulp is kept at a temp. higher than that at which albumins coagulate (drying, however, is avoided) until the outlines of starch particles begin to fade under the microscope. Then it is dried at a temp. lower than that of caramelization and is ground to dust. Of this product 4-6% is added to flours to improve them.

Heat treatment of flour. D W KENT-JONES, C W CHITTY and WOODLANDS, LTD. Brit 338,603, Aug 23, 1929. Various details of app and procedure are described for heating cereal flours for use in ordinary manner or as addns. to untreated flour to improve its properties.

Complex organic peroxides. PILOT LABORATORIES, INC. Brit. 339,336, Aug 20, 1929. Mixts of the acid chlorides of higher fatty acids and benzoyl chloride, chlorobenzoyl chloride or bromobenzoyl chloride are treated with peroxidizing agents. Products are obtained suitable for bleaching flour, oil seed meals, soaps, oils, fats, waxes, egg yolk and other foods products, etc. Acid chlorides derived from the mixed fatty acids of coconut oil may be used as an initial material. Various details of procedure are described.

Bread, etc. JEAN J. PUTSCHER. Fr 37,165, June 28, 1929. Addn to 657,741. Petrolatum oil, glycerol and dextrin are added to the cooking salt ordinarily used in the manuf. of bread, etc., dough.

Irradiating bread or other products in baking ovens. ALBERT WÖRNBR. U. S. 1,790,134, March 10. A Hg-vapor lamp is mounted in a baking oven with a reflector arranged to deflect the rays from the lamp so that all the irradiation is by reflected rays. This is stated to avoid impairment of taste such as direct irradiation may cause.

Use of pectin in baked food products. ARNOLD S. WAIL (to E W Stewart & Co). U. S. 1,795,980, March 10. In making baked goods such as sweet rolls or cake, pectin is added to a sugar soln. used as an ingredient, in order to improve the texture and yield. The proportion of pectin may be about 5% the wt of the flour used.

Pectin. ROGER PAUL and ROBERT H. GRANDSEIGNE. Fr 695,204, Aug 24, 1929. Neutral pectic juices are caused to solidify slowly to a gel at ordinary temp. by adding 2-10% of conc HCl . The gel is afterward finely ground and washed free from HCl by water.

Milk preparation free from milk sugar. JACOB POHLMANN and JACOBUS R. F. RASSERS. Ger. 515,894, Sept 13, 1928. See Brit 320,497 (C. A. 24, 2512).

Apparatus for pasteurizing milk, etc. FRITZ G. CORNELL, JR. (to Jensen Creamery Machinery Co.) U. S. 1,795,932, March 10. Structural features.

Corrugated heat-exchange shell apparatus suitable for pasteurizing milk. EDGAR B. NICHOLS. U. S. 1,797,014, March 17. Structural features.

Tubular heat-exchange apparatus suitable for treating milk, cream, etc. FRITZ G. CORNELL, JR. (to Jensen Creamery Machinery Co.) U. S. 1,795,837, March 10. Structural features.

Apparatus for pasteurizing, aerating and degasifying cream. OLIVER V. JONES and ALVIN SCHNEIDER U S. 1,796,031, March 17.

Apparatus for sterilizing milk. RICHARD SELIGMAN. Ger. 518,462, July 10, 1923. See Brit. 302,743 (C. A. 23, 4279).

Apparatus for spray desiccation of materials such as milk. FRANK H. DOUTHITT (one half to Chester E. Gray) U S. 1,797,055, March 17. Various structural details and details of operation are described. Cf. C. A. 24, 4550.

Conserving animal and vegetable matter, especially foods. RICHARD WILLSTÄTTER. Ger. 515,960, Apr. 16, 1929. Addn. to 513,665 (C. A. 25, 1921). As in 513,665, HCN is used as the conserving agent. In conserving meat, 20 mg. HCN is packed with each kg. of meat.

Preserving food materials such as fruits, vegetables and meats. U. GIOMI. Brit. 338,768, Dec. 17, 1929. Fruits such as strawberries, peaches, grapes or apricots are immersed in pure alc. and then dried before packing. Vegetables and meats may be packed in vessels and then sterilized with steam, and all such products, placed in suitable receptacles, are subjected to a vacuum treatment, treated with an inert gas and then sealed. App. is described.

Preserving foods such as fish or meats. M. T. ZAROSCHENIEFF. Brit. 339,172, Feb. 16, 1929. The materials are rapidly chilled or frozen for preservation or transport by subjection to air satd. with a highly atomized liquid such as fresh or sea water or brine in the form of fog and at a low temp. (suitably about -22° to -25°) and may be subsequently defrosted by treatment with a fog laden atm. at a temp. above freezing. App. and details of procedure are described.

Apparatus for freezing foods such as fish. ALEXANDER H. COOKE (to Atlantic Coast Fisheries Co.) U S. 1,795,330, March 10. Heat-conducting elements depending from an endless-belt carrier serve as heat transfer means between the cooling agent and the articles to be frozen. Various structural details are described.

Meat extracts, etc. BEEFEX, LTD. and J. E. AFS. Brit. 338,808, Jan. 30, 1930. Cubes of material such as meat exts. are toasted to sterilize them and harden their surfaces so that they can be handled without wrapping. Various details of treatment are described.

Vacuum treatment of corn, fruits, etc. A. E. JONSSON. Brit. 338,635, Sept. 9, 1929. Vegetable products are subjected to a vacuum in the presence of a regulated proportion of water which is vaporized by heating during the continuation of the vacuum treatment (a temp. of $30-45^{\circ}$ being suitable for corn).

Apparatus for conditioning raisins or other fruits by dry air treatment. THOMAS W. W. FORRENT (to Sun Maid Raisin Growers of Calif.). U. S. 1,795,302, March 10. Structural features.

Filter press for raisins, etc. PAUL COMBESCURE. Fr. 695,211, Aug. 27, 1929.

Machine for extracting grape juice. KARL MAUREA. Ger. 515,925, Aug. 14, 1928.

Cocoa preparations. G. SEFTNER. Brit. 339,034, Sept. 23, 1929. Radioactive chocolate is prepd., e. g., by admix. with sugars which have been dehydrated and then allowed to cryst. in water contg. emanation or radioactive salts.

Coffee preparations. W. PÄFFGEN. Brit. 339,543, Sept. 18, 1929. Caffeine and 'roasting toxins' are stated to be removed by use of absorption paper impregnated with finely divided active C.

Stock feed. GAGE A. KENT. U S. 1,796,031, March 10. Pea or nut-sized pieces of cracked pressed oil cake are provided with a substantially uniform and continuous coating of sorghum molasses, and each molasses-coated piece is further provided with a final coating of powd. oil-cake meal which prevents adhesion.

Preserving green fodder. SIGMUND VON KAPFF. Ger. 518,463, Oct. 26, 1927. The fodder is sprayed with a soln. contg. HCOONa and a salt or acid capable of liberating HCOOH therefrom.

13—GENERAL INDUSTRIAL CHEMISTRY

HARLAN S. MINER

Good lighting—an executive problem. DEAN M. WARREN. *Ind. Eng. Chem.* 23, 512-6(1931).

The use of emulsions in industry. V. KUBELKA. *Chem. Listy* 24, 354-7, 372-4 (1930), cf. C. A. 24, 5393.—The use of much com. app. for prep. lab. or theoretical emulsions on a com. scale is discussed.

FRANK MARSH

Metafiltration. GEORGES GENIN. *Rev prod chim* 33, 613-6(1930); cf. C. A. 23, 1180; 24, 1061 — This method of filtration is described. P. THOMASANT

The injurious action of smoke, dust, gases, factory fumes, etc., on plants. ANTONIN ROLET. *Rev gén sci* 42, 111-20(1931) — A discussion is given of the origin in the atm of chem gases, vapors from tar roads, illuminating gas and other contaminations, and their effect on plants. J. T. SULLIVAN

Heat transfer in superheaters. W. GUMZ AND F. MICHEL. *Feuerungstech* 18, 129-31, 152-5(1930). T. I. S.

Measurement of air temperatures. M. TOMLINSON. *Heating, Piping and Air Conditioning* 2, 769-70(1930). E. I. S.

Predicting the flow of liquids in pipes. R. E. GOULD. *Heating, Piping and Air Conditioning* 2, 850-2(1930). E. I. S.

Heat transmission to water flowing in pipes. A. E. LAWRENCE AND T. K. SHERWOOD. *Ind Eng Chem* 23, 301-9(1931) — Expts indicate that for heating water in turbulent flow the effect of pipe length on the film coeff of heat transfer is negligible. D. GORDON

Heat transfer through air spaces. P. D. CLOSSE. *Heating Piping and Air Conditioning* 2, 875-8(1930) — A résumé of important investigations. Brief abstracts are presented. T. I. S.

Calculation of loss of head in straight channels. D. KRSTICHER. *Rev. univ. mines* [8] 3, 271-91(1930) — A theoretical mathematical discussion leading to the construction of a logarithmic chart for computation of loss of head in rectilinear channels carrying water, oil, steam or air, account being taken of the viscosity of the fluids. T. I. S.

Method for demustardization of one-lon mustard containers. H. P. AVTRILL AND I. W. LANE. *Chem Warfare* 17, 932-3(1931) — The containers to be demustardized were 2 steel drums which had held mustard for 10 years. 1 or reasons given it was impracticable to use corrosive agents or solvents for mustard to destroy the mustard remaining after emptying the drums. The drums were steamed for a period of 8 hrs with steam at a pressure of 5-15 lb per sq in. This destroyed (hydrolyzed) practically all the mustard, only traces remaining in the approx 75 lb of residual S. A. L. KIMMER

Refrigerants show similarities. J. L. STARR. *Refrigerating Eng* 20, 302-3(1930) — Similarities between 16 different substances were shown in terms of pressure-temp and compressor displacement relations. From a knowledge of one point on the pressure temp curve it is possible from the curves given to predict the pressure exerted by a gas at any given temp. This was done for CCl_2F_2 and exptl. results agreed fairly well with those predicted from the curves. A. H. JOHNSON

Refrigeration and the fishing industry. D. B. FINN. *Refrigerating Eng*, 20, 347-90(1930) — Rapid freezing results in the formation of min ice crystals. Small ice crystals do not rupture the cell walls of the material frozen and therefore when thawed such material more closely resembles the fresh food product than if slow freezing had occurred with consequent formation of large ice crystals. In order not to allow the crystals to grow the storage temp must not be allowed to fluctuate. Also fluctuations of temp set up internal pressures which tend to cause oil to ooze from the material frozen. Loss of moisture from frozen material during storage may be reduced by keeping all objects within the storage chamber at as near the same temp. as possible. In such a case moisture will not condense on colder surfaces such as cooling coils, etc. Sufficient insulation so that the temp differential between the cooled product and the coils is low will prevent serious shrinkage. The occurrence of "rusting," the yellow discoloration which occurs in oily fish during cold storage, may be reduced by using low storage temp.

"Rusting" appears to be concerned with a change in the fat or oil, but nitrogenous constituents and possibly H-ion concn are thought to play an important part. A. H. JOHNSON

Dangerous liquids. B. L. HUNSTON. *Safety Eng* 60, No 3, 174, 176, 178-9(1930) — The main considerations in selection of containers for shipment are, strength and tightness, possible reaction of liquid on containers, and proper handling of container, both loaded and empty. L. I. S.

Role of gas masks in mines. P. ESZRO. *Bányás Kohás Lapok* 63, 55-60(1930) — Modern gas masks (respirators), e. g., the "All-Service-Mask" examd by the U S A Bureau of Mines in May, 1924, yield perfect defense against CO . Therefore gas masks should be used in mines instead of expensive air-regenerating app., as the cost of the former is only 12-15% that of the latter. S. S. DE FINALLY

The moisture in technical gases. III. Examples. FRIEDRICH LÜTH. *Arch*

Eisenhüttenw 4, 281-7(1930); cf. *C. A.* 25, 1924—Numerical examples (10) showing the applications of the method of calcn. are given J. BALOZIAN

Control and regulation of distilling columns. M. PÉREZ *Chimie & Industrie* 25, 286-90(1931)—A brief discussion is given, from a practical operating standpoint, of various points regarding the control and regulation of continuous distg. columns, more particularly columns for distg. alc. A. PAPINFAU COUTURE

Vacuum as an insulator. CHARLES O. DUEVEL, JR. *Refrigerating Eng* 20, 223-8 (1930)—D reviews the subject of heat transfer in vacuum insulation and in gases. An app. is described by the use of which it is possible to measure thermal cond. of vacua and gases A. H. JOHNSON

Separation and size distribution of microscopic particles (ROLLE) 2.

BRADY, GEORGE S. *Materials Handbook*. 2nd ed., revised. New York: McGraw-Hill Book Co., Inc. 600 pp. \$5

Chemiehitte. *Taschenbuch für den praktischen Chemiker*. Edited by AKADEMISCHEN VEREIN HÜTTE, E. V. Berlin: Wilhelm Ernst & Sohn. 914 pp. Linen, M. 28, leather, M. 31

Deutsches Bergbau-Jahrbuch, 1931. 22nd year. *Jahrbuch der deutschen Braunkohlen-Steinkohlen-, Kali und Erzindustrie, der Salinen, des Erdöl- und Asphaltbergbaues*. Edited by H. HIRZ and W. POTTMANN. Halle (Saale): W. Knapp. M. 16

Hütte. *Des Ingenieurs Taschenbuch*. Band I. Grundlagen der Technik. Edited by AKADEMISCHEN VEREIN HÜTTE, E. V. Berlin: Wilhelm Ernst & Sohn. 1199 pp. Linen, M. 17 50, leather, M. 20 50

LEIGH, C. W., and MANGOLD, J. F. *Practical Mechanics and Strength of Materials*. 2nd ed. New York: McGraw-Hill Book Co., Inc. 420 pp. \$2 75. Reviewed in *J. Western Soc. Eng.* 35, 495(1930).

Purifying gases. UNION CHIMIQUE BELGE S. A. *Fr.* 37,418, May 17, 1929. Addn. to 650,516 (*C. A.* 23, 3284). The pptn. of the $(\text{NH}_4)_2\text{SO}_4$ is localized in the part of the app. where the NH_3 is dissolved and is prevented in the part where the cooling takes place. The fall of temp. of the satd. soln. is reduced during the cooling but is afterward repeated several times. Cf. *C. A.* 24, 3843

Separation of gas mixtures, and nitric acid production. N. CARO and A. R. FRANK. *Brit.* 338,497, Aug. 13, 1929. A constituent of a gas or vapor mixt. is sep'd by condensation by cooling under such conditions that the mixt. leaving the condenser is unsatd. with respect to the sep'd constituent. Various details of app. and operation are described. Hot aq. nitrous gases, produced by oxidation of NH_3 , with addn. of steam, H_2O or SO_2 in some cases, can be treated to sep. the water, free from or low in HNO_3 , and if desired the N oxides may be directly worked up to form nitrates or nitrites by reaction with suitable oxides, carbonates, chlorides, etc. Sepn. of water, etc., may be effected by the method described in connection with the production of Ca cyanamide, HCN and NH_3 salts or aq. NH_3 .

Distilling or drying carbonaceous or other solid materials, etc. E. M. SALERNTI and E. M. S. INDUSTRIAL PROCESSES, LTD. *Brit.* 338,939, July 30, 1929. Operations of this general character or evapn. or other heat treatments of liquids are effected in a described app. which may comprise a plurality of juxtaposed troughs through which the material successively passes while at the same time it is subjected to a secondary movement by the action of agitators in a direction counter to that of main passage of the material.

Alumino-thermic reactions. E. BLOCHNER. *Brit.* 338,468, March 11, 1930. In effecting reactions such as those of Fe oxide and Al, very high temps. are attained by employing one or both of the reacting materials in molten state. Various details of app. and procedure are described and the operation may be conducted in a vessel surrounded by N under 200 atm. pressure.

Regulating the temperatures of chemical reactions. HOLZVERKÖHLUNGS-INDUSTRIE A. G. *Brit.* 339,144, Jan. 18, 1929. In regulating temps. such as in the production of acetone from C_2H_2 and water vapor, the reacting constituents are maintained in motion in a continuous circuit past heat-exchanging walls which sep. part of the circuit from a reaction zone, while part of the constituents are diverted continuously into the reaction zone, the reaction temp. is controlled by regulating the speed of travel of the reacting materials in the circuit. App. is described. The procedure is suitable for exothermic or endothermic reactions between liquids or between liquids and gases. *Brit.* 339,148

relates to similar regulation of reactions carried out under pressures above atm. pressure.

Furnace operation for chemical reactions. *SOC. CHIMIQUE DE LA GRANDE PAROISSE (AZOTE ET PRODUITS CHIMIQUES)*. Brit. 338,444, Jan. 28, 1930. A jet of O₂ or of air enriched with O is impinged on a restricted area of a bed of coke to produce a high temp. zone in which refractory materials such as silica, alumina and Ca phosphate can be volatilized and caused to react. Fine powders are sep'd. from the combustion products and in the operation of the process silica or ferric oxide for polishing metals may be reduced to fine powder, and silica or alumina can be caused to react on Ca phosphate for the production of Ca silicate or Ca aluminates and oxides of P. App. is described.

Autolysis or heterolysis of animal or vegetable materials. *SOC. FRANÇAISE DES PRODUITS ALIMENTAIRES AZOTÉS*. Ger. 518,263, Jan. 14, 1925. The operation is effected with the aid of an antiseptic mixt. comprising CHCl₃ and an alc. of lower sp. gr. than the materials under treatment. Thus, fish may be autolyzed with water for 4 days at 40° in the presence of a mixt. of CHCl₃ and BuOH.

Granulating salts. *I. G. FARBENFAB. A.-G.* (Bodo Haak, inventor). Ger. 518,090, Dec. 14, 1927. In granulating salts, particularly fertilizer salts, by causing the fused salt to fall in drops into a liquid, the path of the salt through the liquid is lengthened by stirring the liquid so that it circulates around the vertical axis of the vessel contg. it.

Geistalizing organic liquids such as benzine. *I. G. FARBENFAB. A.-G.* Brit. 339,360, July 4, 1929. To facilitate safe handling and storage, liquids such as benzine or petroleum ether have formed in them a voluminous 'framework' of solid inorganic oxide or hydroxide such as may be produced by the action of a base such as NiH₂ on SnCl₄, SiCl₄ or a chloride of Fe or of Al.

Gel-like benzine. *L. LOENNA*. Hung. 101,509, May 3, 1930. Soap is boiled in alc., then shaken with benzine and water until the mixt. becomes gel-like; e.g., 2 g. soap and 2 g. alc. are boiled, then 500 g. benzine and 2 g. water are gradually added.

Oil or wax emulsion. *L. PÖCS*. Hung. 101,357, Feb. 22, 1930. The aq. phase of the emulsion contains a mol. or a colloidal soln. of halogenated fatty acid soap as emulsifier. A disinfectant agent can be added or the emulsified oil, e.g., turpentine oil, may also have disinfectant qualities.

Od or wax emulsion. *L. PÖCS*. Hung. 101,358, March 5, 1930. Oil or wax is emulsified with soap, and sufficient alcohol or alcoholic KOH is added to give permanent emulsion in water. Finally the disinfectant is added.

Emulsions, ointments, etc. *T. GOLDSCHMIDT A.-G.* Brit. 339,558, March 9, 1929. Emulsions are prep'd. by mixing esters (contg. non-esterified Olf groups) of polyhydric alcs. with higher fatty acids, such as glycerol monostearate or diethylene glycol monostearate, with water or aq. solns. in the presence of acylated alkylene diamines or their deriva. such as the C-substituted or N-alkylated deriva., and various other substances may be emulsified or dispersed in the products, such as glycerol, H₂BO₃, citric acid, fats, fatty acids, waxes, vegetable or mineral oils, rubber, glue, gelatin or dextrin, terpenes, perfumes, latex, cow milk, lampblack, talcum or ZnO. Examples are given. Cf. C. A. 24, 5437.

Lining liquid containers. *STANDFASSENWERKE ROSTOCK & BAERLOCHER and GEORG WALTER*. Austrian 121,533, Oct. 15, 1930. In lining fermentation vats and other liquid containers made of concrete or the like with bituminous, resinous or waxy substances the lining substances are applied as aq. emulsions. Preferably, a relatively stable emulsion is first applied, followed by a relatively unstable emulsion. Fibrous materials and other fillers may be included in the emulsions, and the addn. of a small quantity of a hydraulic mortar is advantageous.

Detecting leakage of noxious gases from pipes. *GABRIELE WINOISCH*. Austrian 121,257, Sept. 15, 1930. The pipes conveying the gases are enclosed in protective pipes, which are connected through a manometer to a suction app. Means for sealing off sections of the protective piping may be provided to assist in locating a leak.

Use of substances of low-boiling point such as carbon dioxide, ammonia or methyl ether for operating engines. *OLGA, FÄHNCHEN ZUR LIPPE and M. BRAND*. Brit. 338,894, June 28, 1929.

Use of hydrochloric acid to prepare bore holes for blasting. *V. TSCHERN*. Brit. 338,454, Feb. 11, 1929. Bore holes are provided with spherical enlargements to receive blasting charges by forcing acid such as HCl in fine jets through the nozzle of a described app., and a preliminary charge may be exploded in the bore hole to facilitate the action of the acid.

Drying air surrounding perishable goods during shipment. *T. RONNIGS, F. C.*

DFM FR and SILICA GEL, LTD. Brit 334,855, July 4, 1929. The air supplied to containers in cars, ships, etc., is dried by silica gel or like material. App is described.

Refrigerating chemicals. N CARO and A R. FRANK. Brit 339,194, Feb 8, 1929. In processes for absorbing at low temps., liquefying or freezing out N oxides, HNO_3 and their mixts., an aq. NH_3 soln is used as the "cold-conveying" medium and substances such as NH_4NO_3 , metal nitrates and metal thiocyanates may be added to the soln. Various details of concns. of solns for different temps. are given.

Refrigerator boxes formed of Celotex or the like dipped in asphalt. J C STATON (to Coca Cola Co.) Brit 338,959-60, Jan 9, 1929.

Refrigerating apparatus using solid carbon dioxide. W T HEDLUND (to Electro-lux Ltd.) Brit 338,795, Jan 12, 1929. Structural features.

Refrigerating apparatus utilizing solid carbon dioxide. WALTER S JOSEPHSON and THOMAS B SLATE (to Dryice Equipment Corp. and American Patents Development Corp.) U S 1,796,907, March 17. U S 1,796,908-9-10 (THOMAS B SLATE to American Patents Development Corp.) also relate to app for the same general purpose. Cf () 24, 446, 2517.

Heat-insulating material suitable for building construction. GEORGE B STRYKER, JR (to Weatherproof Products Co.) U S 1,796,631, March 17. A base of compressed fibrous material such as compressed waste paper and cotton fiber is provided with a top layer consisting of a hardened mixt. of fibrous material with cement, clay and lime. Cf () 24, 446, 2517.

Heat insulators. S A I G (Soc. Anon. Frenzioli, Guadagnini) Brit. 338,743, Dec 3, 1929. See Fr 686,226 (C A 25, 549).

Insulating electric wires and cables. FRANCIS J BRISLEE (to British Insulated Cables, Ltd.) U S 1,795,994, March 10. An esterified cellulose compn. such as a cellulose acetate compn. in hard and substantially solvent free form is applied in layers with textile supports around a conductor under approx. atm. pressure, and the material is then heated to about 100° , at which temp. the compn. softens and flows so that adjacent layers unite and impregnation of the textile supporting material is effected.

Insulation of wire for electrical purposes. I G FARBENING A-G. Ger 518,328, April 24, 1929. The insulation consists of filaments of butadiene hydrocarbon polymerization or condensation products which have been hardened so that they no longer show the extensibility and elasticity of rubber. Hardening may be effected by heat or the action of sulfurizing or oxidizing agents. Examples are given.

14—WATER, SEWAGE AND SANITATION

EDWARD BARTOW

Water supply work of the U. S. Public Health Service. ARTHUR P. MILLER. *Water Works Eng.* 83, 907-8(1930).—The relations of the P. H. S. officials with local health agencies, water and sewage treatment plant operators and the like are usually of an advisory character. A review of the manner of certifying water supplies to be used on common carriers engaged in interstate traffic is given. The objectionable practice of adding ice directly to the drinking water used on railway coaches has been nearly eliminated. The effect on the drinking water by careless handling by railway employees is being studied. Since the beginning of the certification of water supplies, the P. H. S. has assumed the entire responsibility of assuring water of good quality on vessels. In 1921 the P. H. S. arranged to do the sanitary engineering work for the National Park Service in the national parks. The work included development of water supplies, building of sewage treatment plants, mosquito eradication and disposal of combustible and non combustible waste.

Study of a spring in Monferrato (Italy). GIOVANNI CALVI. *Giorn. farm. chim.* 79, 543-8(1930).—A qual and quant. chem. analysis was carried out on the water of an inexhaustible spring. G SCHWOCH

New pipe materials for Dresden water mains. SCHEMEL. *Gas u. Wasserfach* 74, 169-73(1931).—Various types of Fe and steel water mains with and without internal coatings, as well as reinforced concrete and wooden mains, are discussed. R W R.

Piping materials for mineral water. E. MAURER. *Gesundh.-Ing.* 53, 333(1930). *Wasser u. Abwasser* 27, 369-70.—Most mineral waters in themselves contain no substances corrosive to Fe or other piping. By far the most important agent in corrosion is O_2 , introduced from the air. O_2 is not normally present in mineral waters obtained from deep in the earth. CO_2 has little or no corrosive effect. Incrustants will deposit on any

pipe but in less quantity where O_2 is absent. Hence to reduce both corrosion and incrustation, the piping system should be planned to avoid entry of O_2 . Wood and galvanized Fe are shortlived while Sn, Pb and Cu pipes are too expensive. Cast Fe is the best all around pipe for mineral waters. C R FELLERS

Protection of water supply pipe lines system from electrolysis with particular reference to insulation, pipe bonding, electrical drainage and return systems. J. J. LAUDIG, et al. *Am Ry Eng Assoc Proc* 32, 415-9(1931), Bull 333.—Regular inspections with tests are recommended where electrolysis is suspected and correction is made by insulation or bonding elec drains where needed. R C BARDWELL

Taste and odor troubles in the Minneapolis water supply. F RAAB, J. *Am Water Works Assoc* 23, 430-4(1931).—Certain types of algae, growing in the filters, were the producers of varying and unpleasant tastes and odors. Alum, chlorine and ammonia were tried without effect. Prechlorination, however, was very successful and increased noticeably the filter runs. D K FRENCH

Elimination of taste and odor in the water supply of Lancaster, Pennsylvania. EDWARD D RUTH, J. *Am Water Works Assoc* 23, 390-9(1931).—Tastes and odors of various kinds were always present. The filter plant was overloaded. Much of the time the water was not safe to drink. By using anhyd ammonia and chlorine, almost perfect results were obtained. D K FRENCH

Successful superchlorination and dechlorination for medicinal taste of a well supply, Jamaica, N. Y. FRANK E HALE, J. *Am Water Works Assoc* 23, 373-86(1931).—The water contained Fe and Mn, and the growth of *Crenothrix* was quite noticeable. Treatment with Cl as required produced a medicinal taste. Investigation showed that leaking gasoline, rather than algae, was responsible. After much exptl. work it was found that superchlorination followed by dechlorination gave the best results. SO_2 was used. The method was inexpensive and quickly applied. D K FRENCH

Interpretation of water analysis. Equilibrium considerations, determining activities and concentration of ions. D S MCKINNEY, *Ind Eng Chem, Anal Ed* 3, 192-7(1931).—Present methods of detg the anions of the slightly dissociated acids (carbonic and phosphoric and their salts) are stoichiometric, and disregard the equilibria which control the relations of these anions and the H and OH ions and hence give the operator a false impression of the quantity of each anion present. Further, the results obtained by present methods cannot be evaluated when more than one weak acid is present, unless unwarranted assumptions are made. Derivation and a list of formulas for the calcns. of the fractional activity or activity ratios of various ions obtained from H_2CO_3 and H_2PO_4 at 25° are given. In order to obtain the ion concn and activities, it is necessary to know the activity coeffs. The relation between the activity and the concns may be expressed by means of the simplified equation of Debye and Hückel. Methods of application of the equation to actual water analysis are given and explained. WAYNE L. DENMAN

An investigation into the clogging of the filter beds at Topchanchi waterworks during hot weather. B K. MANDAL, *Indian Med. Gaz* 66, 84-5(1931).—The growth as well as the actual decay of the lower forms of vegetable life (particularly *Crenothrix*) are the more important factors influencing the clogging of filter beds during hot weather. An increase in temp with an acceleration in the production of org. matter cause a simultaneous reduction in the dissolved O_2 content, and the evolution of CO_2 , these conditions are favorable for the production of the so-called "iron bacterium" and allied forms of life. FREDERICK G GERMUTH

Chemical control and general supervision of water-treating plants. R. M. STIMMEL, et al. *Am Ry. Eng. Proc.* 32, 419-22(1931), Bull. 333.—Survey indicates field inspection is desirable in addn to lab check, the field-test method for rapid boiler water analysis recommended in A. R. E. A. Manual is used. Treating plant results should be coordinated with the performance of water in locomotive boilers. R. C. BARDWELL

Treats water twice in new plant. R. V. ZAHM, *Ry Eng and Maintenance* 27, 363-4, 377(1931).—The Missouri, Kansas and Texas R. R. installed a complete 30,000 gal per hr water-softening plant at its Franklin, Mo., terminal in 1929. Lime, soda ash and Na aluminate are used to soften and clarify a reservoir water ranging from 5.0 to 7.5 gr. per gal total hardness and 4.0 to 20.0 gr per gal. suspended matter. Addnl. treatment with alum to neutralize OH and Na hypochlorite for sterilization is given that portion used for drinking purposes with subsequent filtration. A Sparling meter is used for lime and soda ash control. Aluminate is added separately by an automatic dry feeder. A min. settling time of 6.8 hrs with a max. of 13.8 hrs. clarifies the water satisfactorily. Hardness of treated water averages consistently under 1.5 gr. per gal. R C B

Results of water treatment. T. F. POWERS *Ry. Age* 90, 329-31(1930).—In 1911 the Chicago and Northwestern RR had a leaky flue failure each 58,633 locomotive miles but with supervised water treatment this had been reduced by 1929 to one failure per 4,343,302. Corrosion and pitting still constitute a problem and embrittlement is a potential danger. A properly supervised wayside treating plant has been found a most satisfactory method for correcting water quality. R. C. HARDWELL.

Lump or hydrated lime? O. T. REES, *et al.* *Ry. Eng. and Maintenance* 27, 276-8, 385(1931).—Available CaO for use in water treatment can be obtained in lump lime at as much as 55% lower first cost than in hydrated lime. However, rapid deterioration of lump lime in shipment and storage affects the reliability and where more uniform results and greater accuracy of treatment are required together with greater ease of control, use of hydrated lime is favored. R. C. HARDWELL.

Experiments with chlorinated copperas as a coagulant. A. CLINTON DECKER, *Water Works Eng.* 83, 887-8(1930), *cf C A* 23, 6525, 24, 1447, 2818.—The source of supply for the Chickasaw, Ala., filtration plant originates in springs and flows through a heavily cypress wooded territory which imparts to the water a color ranging from 40 to 130 p.p.m. Alky., total hardness and turbidity are low. Color removal was complete when 0.7 grain per gal. of copperas, mixed with 2.16 p.p.m. Cl_2 , was added as a coagulant. The Cl_2 added was sufficient to oxidize the Fe and also to satisfy the Cl_2 demand of the raw water. The filtered water is adjusted with lime to a pH of 7.0-7.2. The increase in color is less when the lime is introduced into the clear water well than when introduced into the coagulating basin. The length of the filter runs has been increased by this process as compared with alum and lime, thus lessening by 50% or more the wash water required. The cost of chemicals is approx \$10 less per million gal. water. C. H. B.

Classification of water movement (phenomena). A. REUTHER, *Gesundh.-Ing.* 53, 164-7, 180-5, 194-202(1930), *Wasser u. Abwasser* 27, 163.—A math. discussion is given with formulas. C. R. FELLERS.

Economics of reservoirs. PAUL HANSEN, *Water Works Eng.* 83, 943-4, 982-6(1930).—Impounding reservoirs should be formed which will meet the requirements for water during the dryest period likely to occur within the period of anticipated demands. Because of the cost of constructing dams and spillways and the difficulty of bonding new work on dams with old it is not usually economical to provide impounded storage for less than 10 yrs and often provision for 20 to 30 yrs is justified. There should be sufficient filtered water storage to neutralize the differential (for ordinary domestic purposes $\frac{1}{4}$ to $\frac{1}{2}$ the total day's consumption is usually required) between the uniform rate of filtration and the variable rate of consumption, plus a suitable reserve for fires and that needed for washing filters, less elevated storage on the distribution system. Cost of elevated storage should be balanced against cost of pumping. Steel is considered best for tanks built on towers. At present, 2 million gal. is the max. practicable size for such steel tanks. For the construction of the smaller sizes of exposed reservoirs built on top of the ground steel is cheaper than reinforced concrete. For covered reservoirs steel has a small advantage in first cost with capacities up to 5 million gal. Tables showing the comparative costs of constructing different types of reservoirs are given. C. H. B.

Overcoming algae troubles in a clear-water reservoir. V. BERNARD SIEMS, *Water Works Eng.* 83, 739-40(1930).—The water supply at Greenville, Pa., is obtained from a stream on which are located 2 reservoirs in series with a sedimentation basin at the lower end. The water then flows to rapid sand gravity filters and Cl_2 is added to the effluent which passes to the clear water reservoir. Doses of 2 p.p.m. CuSO_4 to the clear water reservoir cut off for several days, followed by doses of 1 p.p.m. CuSO_4 to the sedimentation basin, did not satisfactorily prevent algae, especially in hot weather. At present, 1.0 p.p.m. NH_4OH is added between the lime and the alum dosages. The Cl_2 dosage is 1.5 p.p.m. which gives a residual Cl_2 of 0.2-0.3 p.p.m. in the distribution system. Results obtained during 2 months' operation indicate that more Cl_2 can be added without causing tastes and odors, as was formerly the case. This permits a higher factor of safety in sterilization. In addn. Cl_2 is more effective. C. H. B.

Turbidity, plankton and mineral content of Detroit water supply. BERT HUDGINS, *J. Am. Water Works Assoc.* 23, 435-44(1931).—As a result of the shallowness of Lake St. Clair, the water could readily become turbid and carry plant and animal life. Several types which give the most trouble in water were identified. At the present time little trouble is being experienced and the present plant is considered satisfactory. The natural conditions which are combating the growth of objectionable forms are discussed. D. K. FRENCH.

Water softening for small municipalities. C. P. HOOVER, *Am. City* 42, No. 3, 7(1930), *cf C A* 23, 656; 24, 4106.—Most of the 322 incorporated communities in

Ohio with populations of less than 10,000 use well water for their municipal water supplies. Many such supplies are hard and contain considerable Fe. A hardness of 90 p p m, or less seems to be acceptable in this section. High per capita cost together with limited and mediocre technical supervision makes the designing of water-softening plants in small communities difficult. Carbonation has overcome the objections to the taste of softened water and has also made it possible to distribute softened well water without filtration.

C. H. BADGER

Water softening: Some properties of certain base-exchange materials. L. AUSTIN R. MARTIN *J Soc Chem Ind.* 49, 389-94T(1930); cf *C A* 24, 1688—A discussion is given of the bulk ds and vol of interstitial spaces, exchange values at various rates of flow of water, regeneration, mech disintegration and losses and contamination of water with SiO_2 , of Basex, Doucil, Borromite, Kenzelite, Zepholite, Refinite, Natrolith and Nat-Rol. There is greatest exchange value and mech disintegration in the synthetic material.

O. M. SMITH

Effect of the disposal of water-softening-plant sludge through the sewage-disposal plant. E. F. ELDRIDGE *Mich Eng Expt Sta., Bull* 34, 7 pp (1931)—Twelve 1500 samples of fresh solids were placed in individual 3-l bottles which were connected to inverted glass cylinders for gas collection and measurement. Two bottles contained no added softening plant sludge, 2 were treated with a dry wt of lime-softening-plant sludge equiv to 25% of the total solids in the sewage sludge, other samples contained 50, 75, 100 and 150%, resp, of the softening-plant sludge. The total wts of sewage solids in all of the samples were practically identical at 41.6-42.6 g. Max gas production, 5.2 l, was obtained from the samples contg 50% added softening-plant sludge. The optimum p_H for sludge digestion is approx 7.2-7.6. Excessive amts of the lime sludge produced an unfavorable p_H for sludge digestion. The expts are being continued.

C. R. FELLERS

Concentration control of boiler water. HERBERT S. WHITON *Power Plant Eng* 35, 452-4(1931)—Control of concn is necessary to avoid foaming and steam contaminated with impurities, sol or insol, from the boiler. In certain cases the Baumé hydrometer, especially calibrated, can be used. In others, the titration of chloride as a measure of total concn is valuable. Cond methods also have a place. Frequent chem. analyses may be necessary. The concn by products of various methods of treatment are described. Blowing down the boiler will reduce its concn and various methods for its regulation are given.

D. K. FRENCH

Relative cost of eliminating impurities in locomotive boiler waters and the value of treatment with respect to chemicals and compounds applied direct to locomotive boilers and roadside tanks, and conditions under which they may be desirable. C. H. KOVL, et al *Am Ry Eng Assoc Proc* 32, 402-5(1931), *Bull* 333.—The phys. condition of the sludge formed is important in the successful internal treatment of locomotive boiler waters. Chemicals forming large heavy particles are desirable. A limit of 15 gr per gal alkali salts is recommended. Good results can be obtained with soda ash or proprietary compds under careful supervision, the cost varying with the quality of the water.

R. C. BARDWELL

Boiler-tube corrosion halted by hot-process treatment. R. E. COUGHLIN. *Ry. Age* 89, 855-6(1930); *Ry Eng and Maintenance* 26, 550(1930)—The Chicago and Northwestern RR eliminated serious pitting in six 500 h p, water tube boilers using Chicago city water, by installation in 1928 of a hot-process softener using lime, soda ash and Na aluminate, this replaced unsatisfactory internal treatment.

R. C. B.

Experiments and results with boiler-water recycling. FEIGE AND WEISS *Arch Wärmewirt* 12, 16-8(1931)—A rather full account is given of practical expts on a 1000 h p boiler in returning blow-down to the feed water continuously through a 3 mm. orifice. No difficulty was experienced, and the indicated saving was 1500 M. per year.

ERNEST W. THIELE

Water requirements in the textile industry. KEHREN AND STOMMEL. *Z ges Textilind* 32, 651-3, 676-7, 698-700, 798-800, 811-4, 830-1, 845-7, 860-2(1929); *Wasser u Abwasser* 27, 170—Ground waters and their suitability for use in dye factories, bleaching plants, laundries and creameries are discussed, together with methods of softening, Fe and Mn removal, filtration, Al coagulation and decolorization.

C. R. FELLERS

Past and present developments in sewage disposal and purification. H. W. CLARK. *Sewage Works J.* 2, 561-71(1930).—A general discussion touching on all phases of sewage disposal with particular reference to the developments at the Lawrence Expt. Sta.

E. HURWITZ

Review of present sewage-disposal technic. HERMANN BACH. *Vom Wasser* 4,

160-86(1930). *U S Pub Health Eng Abstracts* 11, S, 25(Feb 28, 1931)—A comprehensive review C. R. FELLERS

Conditions affecting general layout of sewage-treatment works. PAUL HANSEN *Sewage Works J* 2, 572-81(1930)—In the design of sewage treatment works provision must be made for population growths and changes in character of sewage. Each element of the works should be so arranged that future extension can be made continuously to the initial installation for economy and improved operating control E. HURWITZ

Operation and control of sewage-treatment plants. CHAS. C. AGAR. *Sewage Works J* 2, 597-614(1930)—A report of the Division of Sanitation, New York Dept. of Health contg suggestions on operation and control of all units which might be embodied in a treatment plant E. HURWITZ

Laboratory service in sewage treatment. M. STAAR NICHOLS. *Sewage Works J* 2, 591-6(1930)—N proposed classifying lab service in 4 divisions: (1) operator's lab, (2) field-control lab, (3) central or state lab, and (4) central or field research lab. The duties of each division are specified E. HURWITZ

Sewage filter media and loadings. WM F STANLEY. *Sewage Works J* 2, 480-94(1930)—See *C A* 25, 1928 E. HURWITZ

Chert gravel as sewage filter stone. J E LAMAR. *Sewage Works J* 2, 495-9(1930)—Natural chert stream gravel gave satisfactory results in that part of the filter more than 3 feet from the surface of the bed. The gravel supported biol growth and was not materially affected by soln. Chipping at the surface of the bed indicated that under long-continued use it might in time have caused "pooling" E. HURWITZ

Use of domestic sewage purifiers. L DE FRÁLY. *Technika* 11, 20-6(1930)—A general review is given of different systems, the principles of domestic purifiers based on artificial biol oxidation methods are discussed, and two newly erected purifiers with activated sludge treatment are described S. S. DE FRÁLY

Observations on acum in the Imhoff tanks of the Ruhr District. KAL IMHOFF. *Sewage Works J* 2, 549-54(1930)—See *C A* 25, 369 E. HURWITZ

Effect of gases on growth of bacteria. R G URTON. *Proc. 12th Texas Water-Works Short School* 1930, 151-3. *U S Pub Health Eng Abstracts* 11, S, 23(Feb 28, 1931), cf *C A* 24, 184—In order to test the effect of partially vacuumizing Imhoff tanks to remove the gases of decompn, lab expts with flasks contg sewage were partially vacuumized and the effect on bacterial nos was noted. In general, there was a slight increase in nos by the agar plate method where the samples were held under 15-8 Hg vacuum. There was no significant difference in the nos. of *H₂S* formers or anaerobes. Conclusion: The toxic effect of sewer gas on bacteria and other organisms that might aid in decompn of sewage sludge in the digestion chambers is slight or negligible C. R. FELLERS

Activated-sludge experiments at the Calumet Sewage Treatment Works. F. W. MOHLMAN AND C E WHEELER. *Sewage Works J* 2, 529-48(1930)—The paddle-wheel diffused air process of activated sludge shows considerable promise in reducing power requirements. The scheme is readily adaptable to large units and is flexible in that requirements for agitation and for O may be adjusted separately so that the air rate need not be excessive merely to maintain the required degree of agitation. The tests indicated that 10 to 15 h p. per million g. is sufficient to produce a satisfactory effluent. With the 2 stage process for activated sludge good results were obtained with small quantities of air. However, this process is very sensitive biologically and requires careful control. Bulking of sludge occurred occasionally in the first stage. The sludge produced by the 2nd stage was generally well flocculated and readily settleable. E. HURWITZ

Effect of iron on the anaerobic decomposition of sewage sludge. L. R. SETTER. *Sewage Works J* 2, 504-20(1930)—Gasification of sewage sludge is not affected by $FeCl_3$ in amounts up to 5 p. p. m. Amounts of $FeCl_3$ from 5 to 20 p. p. m. cause a progressive retardation of gas formation with consequent doubling or tripling of the digestion period. Retardation was partly eliminated by neutralization of the acidity caused by the hydrolysis of $FeCl_3$. Addn of $FeCl_3$ up to 10 p. p. m. was not toxic to the organisms of sewage sludge, greater amounts showed decided reductions of organisms and amounts as high as 20 p. p. m. completely reduced the microbial population with the exception of the sewage fungi. Iron added as the element, oxide or hydroxide had no appreciable effect on digestion, the acetate, citrate, $FeCl_3$, and $FeSO_4$ caused slight retardation. E. HURWITZ

Rejuvenated sludge. MAX PRÜSS. *Sewage Works J* 2, 477-85(1930)—Sludge can be rejuvenated by daily vertical mixing of the fresh solids with all the old sludge in a digestion tank. Vertical mixing with the app described is so thorough that there are

no fresh sludge pockets and fly and odor nuisance due to pumping fresh sludge on to the drying beds is eliminated. Mixing increases the gas yield 100%, reduces the org. content of the sludge, breaks up scum formation and tends to equalize the pH value and maintain a uniform temp throughout the tank. E. FLURWITZ

Correlation between biochemical oxygen demand and suspended solids of activated-sludge effluent. H. GLADYS SWORE. *Sewage Works J.* 2, 500-3(1930) —Results of expts show that the biochem O demand of activated-sludge effluents increases with an increase in suspended solids. The relative increase in biochem O demand is greater as the period of incubation is lengthened. E. FLURWITZ

Biological oxidation of sulfur. IV. Influence on ammonification and nitrification in activated sludge. C. V. R. AYYAR. *J. Indian Inst. Sci.* 13A, 165-71(1930) —Suspensions of sulfur and activated sludge, constantly aerated with ammonia free air, showed a rapid fall in ammonia during the first 3 days with a corresponding decrease in bacterial numbers. But ammonia was rapidly produced from the 11th to 18th day and large numbers of fungi appeared. After the 10th day, nitrates accumulated steadily until acidity developed to about pH 5.4. With phosphates present, the active nitrification proceeded longer before being hindered by the acidity. Development of acidity pH 2.4 reduced the ammonia somewhat in the second 3 weeks. A sudden decrease in ammonia from the 20th to the 22nd day may have been caused by fungus activity. The steady increase in ammonia after the 22nd day may have been due to the colloids, flocculated by the H_2SO_4 , leaching out the adsorbed ammonia. With phosphate present, the changes in the second 3 weeks were similar but less intense. M. E. L.

Making and interpreting a sanitary survey. EDWIN C. FURD. *Sewage Works J.* 2, 582-90(1930) —The manner of making a survey, the elements to be considered, and the method of obtaining desirable information are considered. E. FLURWITZ

Recent progress in air conditioning. W. H. CARRIER. *Refrigerating Eng.* 21, 187-9(1931) —Air conditioning is discussed in relation to the comfort of the individual, fields of application and mech. and engineering requirements. A. H. JOHNSON

Air conditioning on passenger equipment. R. W. WATERFILL. *Ry. Age* 90, 233-4(1931) —Air-conditioning equipment being tested in passenger car service, weighing approx 5 tons, operates by thermostatic control supplemented by hand dampers regulating raw outdoor air. The refrigerating app. uses Carrene (CH_2Cl_2) and Triline ($C_2H_2Cl_4$) which are non inflammable and safe. Practical air conditioning in railway cars is considered to be in the early development stage. R. C. BARDWELL

Soot particles in New York City air. E. E. FREE. *Fuels and Steam Power* (A.S.M.E. Trans.) 53, 9-12(1931). —Counts were made by an impact dust counter. The more important variable affecting results was the direction of the wind. Another is clumping of the soot particles, the cause of which has not been detd. An av. of approx. a half million dust and soot particles is to be found in each cu. ft. of N. Y. City air. At least 90% is ordinary soot or unburned C. The following were also found, auto, tire rubber, asphalt, brick, rock, org. materials, slivers of glass, glazed tile, leather, cotton fiber, hair and a few living mold spores. The size varies from 0.0001 to 0.0500 in. At 100 ft. above the street level the no. of soot particles decreases 25 to 50% at 5th Avenue and 23rd St.; no high level counts were made at other points. For the air over the entire city the av. amt. of suspended matter is 4 to 5 short tons. The relative absence of ash indicates that most of the city's soot and air dust comes from small fires in apartment houses, office buildings and homes, rather than from smoke stacks of the large power houses or other industrial plants. H. C. PARISH

Investigation of atmospheric pollution—report of superintendent of observations. J. S. OWENS, et al. *Dept. Sci. Ind. Research 16th Rept.* 1931, 7-26 —The rept. contains a list of stations, results obtained with the deposit gage, results obtained with the automatic filter, measurement of daylight and deposit of microorganisms at Stockport. Tables show the mean monthly deposits recorded in 1929-30, classification of mean monthly deposits recorded in 1929-30, expressed as % of the general av. for the respective stations and hourly variations of suspended impurity, in different seasons. Eighty-four stations have deposit gages and 11 have automatic filters. No conclusions are drawn. W. H. BOYNTON

Stream pollution from the operator's point of view. EARLE B. PHELPS. *Sewage Works J.* 2, 555-60(1930) —There are 3 types of stream pollution with which the operator of a sewage plant has to contend, namely, phys., chem. and bacteriol. Sewage-treatment works are designed to deal with each of these types independently or as a whole depending on the capacity and limitation of the stream into which the effluent is to be discharged and on the uses to which the stream is to be put. A knowledge of conditions down stream facilitates the intelligent operation of the plant. E. H.

Importance of the vacuum system of sludge removal and septic-tank cleaning. A. RINGEL *Gesundh.-Ing* 1930, 308-14, *Wasser u Abwasser* 27, 176 —The advantages of the method are outlined
C. R. FELLERS

Removal of ordinary sewer gases. A SALMONY-KARSTEN *Wasser u Gas* 1930, 785-7, *Wasser u Abwasser* 27, 177, cf C. A. 25, 1313 —A patented app. is described
C. R. FELLERS

Packing-plant waste treatment. M. STARR NICHOLS AND JOHN C. MACKIN *Sewage Works J* 2, 435-42 (1930) —Treatment of packing house waste by plain sedimentation and sprinkling filter resulted in 90% purification based on removal of suspended solids, org. N and fats and about 80% based on reduction in biochem. O demand
I. HURWITZ

Use of sewage gas as city gas (FULWIELE) 21. Leather, sanitation and colloid chemistry (WILSON) 29. Mottled enamel (McKAY) 11G. Filter [for water] (Ger. pat. 515,852) 1.

PLANCHON, RENÉ *Traitement industriel et rationnel des sous-produits d'abattoirs et des déchets organiques*. Paris Dunod 460 pp. F. 140, bound, F. 150

Filter suitable for filtering water through sand. STANISLAUS J. RICKS and CHARLES E. DOUGLAS. U. S. 1,796,960, March 17. Structural features

Apparatus for softening water with base-exchange materials. S. A. KRIGER Brit. 339,510, Dec. 21, 1929. Structural features

Transportable plant for sterilizing water by chlorination. JOSEF MUCHKA. Austrian 121,030, Aug. 15, 1930

Apparatus for addition of corrosion-preventive chemicals to water in hot-water systems. HENRY L. SHULDBERGER. U. S. 1,796,407, March 17. Structural features

Electric system for protecting boilers. ALEXANDER KIRKALDY (to Electro Anti-Corrosion Corp.). U. S. 1,796,715, March 17

Device for degassing boiler-plant water. MASCHINENFABRIK OERLIKON. Ger. 515,597, Mar. 27, 1928

Boiler-scale remover. Z. V. PIVNEV and S. N. IGNATCHENKO. Russ. appl. 44,852, Apr. 12, 1929. A mixt. of K_2CO_3 and $CuSO_4$ is placed in the boiler. The mixt. is then removed with water together with the scale after a certain period

Use of sodium aluminate to increase the rate of anaerobic digestion of sewage solids. WILHELM RUDOLFS (to National Aluminate Corp.). U. S. 1,797,157, March 17.

Rotary sewage distributor. E. HARTLEY. Brit. 338,650, Sept. 20, 1929. Structural features

Aeration tank (with a power-driven agitator) for sewage treatment by activated sludge. KARL IMHOFF, FRANZ FRIES and FRIEDRICH SIERP (Fries and Sierp to Imhoff). U. S. 1,797,147, March 17. Structural features

Apparatus for incinerating garbage, etc. ALEXANDER A. GOLOVITCHIKOV (one-half to Shunichi Ono). U. S. 1,795,771, March 10. Structural features

Apparatus for clarifying waste waters of the paper, cellulose, etc., industries. ERNST BUCHHAAS. Ger. 518,063, June 5, 1928. See Austrian 114,184 (C. A. 24, 189).

15—SOILS, FERTILIZERS AND AGRICULTURAL POISONS

J. J. SKINNER AND M. S. ANDERSON

Method and procedure of soil analysis used in the Division of Soil Chemistry and Physics. W. O. ROBINSON. U. S. Dept. Agr., Circ. 139, 1-19 (1930). —The fusion method of soil analysis developed for use in the soil chemistry lab. of the Bureau of Chemistry and Soils is described. Directions are given for the prepn. of the soil sample after it reaches the lab. The limits of error allowable in a satisfactory analysis are discussed and the limits of error for various elements are given in percentages and in lbs. per acre. Methods and procedures are given for the detns. of moisture, loss on ignition, org. matter, Cl, F and the oxides of Si, Ti, Al, Fe, Mn, Ca, Mg, K, Na, P and S. The method is designed to show accurately the ultimate comp. of the soil for scientific purposes, but it is claimed to have only a very general application to problems involved in soil fertility studies.
W. H. ROSS

The study of the dynamics of the absorbing complex of soils. M. A. VINOKUROV. *Pochvoedenie* (Pedology) 23, No. 3-4, 46-92 (in English 92-3) (1928) —In a series of

expts. on the A and B horizons of a chernozem soil the amt. of exchangeable Ca and Mg. exchange capacity and unsatn. were detd. during the vegetation period on different systems of cropping, and during the various seasons. It was noted that the Ca was higher in the fall and lower in the spring. The reverse is true for Mg. The total bases was lower in the spring. The base-exchange capacity as detd. by the BaCl₂ method showed some variations. During the first part of the growing season the total exchange capacity decreased slightly, whereas the Ca increased. In the spring the soil was found to be slightly unsatd. In the B horizon the unsatn. was greater in the spring, and in A it was greater in the fall. In some cases, especially on the soils in sod the total Ca and Mg absorbed was greater than the base-exchange capacity. The suggestion is made to consider the absorption capacity as detd. by the Ba method as the exchange or true absorption capacity. The increase in absorption capacity is also to be considered as absorption capacity, but the cations are not exchangeable. During the summer and winter seasons there is an increase in absorption capacity. Vinokurov ascribes this increase to the changes in pH and perhaps mech. absorption. The quantity of highly dispersed particles of the absorption complex drops beginning with the spring season up toward the middle of summer. It is due to the state of satn. of the complex, it is more unsatd. in the spring. J. S. JOFFE

Experiences with the Neubauer method for determining mineral nutrient deficiencies in soils. S. F. THORNTON. *J. Am. Soc. Agron.* 23, 195-208(1931).—Comparisons are made of the 0.2 N HNO₃ extn., Neubauer method, Illinois phosphate test, Hoffer stalk test, pot tests and field yields on 6 soil series. The Neubauer method gives results which agree most closely with those of pot and field tests. The 0.2 N HNO₃ extn. and Illinois phosphate tests give extremely high P₂O₅ results for all soils having received applications of phosphate rock and for soils of the Culver sand series regardless of previous fertilization. The Neubauer method sets as tentative limits for deficiencies for field crops under Indiana farming conditions 4 mg. P₂O₅ and 10 mg. K₂O. Nutrient absorption by seedlings is greatly affected by the selection of seed and temp. control and in a minor way by light intensity, moisture content, soil reaction and the presence of other nutrients. J. R. ADAMS

A method of oxidizing and dissolving soil for the determination of total and filtrable manganese and phosphorus. E. M. ENCKERT. *Soil Science* 31, 175-82(1931).—Soils are prepd. for total Mn and P analyses by the use of H₂SO₄ and NaClO₄ to destroy org. matter and to bring minerals into soln. One g. of soil is placed in a Kjeldahl flask, 2 g. NaClO₄ added, then 25 cc. of 50% (by vol.) H₂SO₄, and the flask is gently heated. If the acid concn. is correct and the heat applied properly, ClO₄ will not accumulate in sufficient quantity to cause an explosion. P detd. in solns. prepd. in this way is greater than when the Mg(NO₃)₂ method is used. A method for detg. available Mn and P consists of shaking 200 g. of soil with 400 cc. of H₂O, filtering through a Whatman No. 2 filter paper and making detns. in the filtrate, which may contain some colloidal material. It is assumed that the soln. contains the approx. amt. of Mg and P present in available form. M. S. ANDERSON

Profile studies in the western province with reference to hardpan formation. M. S. DU TOIT AND J. REYNERS. *South African J. Sci.* 27, 280-95(1930).—Complete chem. analyses are given for soils of 3 profiles. There appear to be 2 extreme modes of hardpan formation: one involves accumulations of SiO₂ under and alk. conditions and in some cases subsequent dehydration and building up of complex silicates, while the other mode consists of accumulation of sesquioxides and in some cases humus or both. Later desiccation of this may result in the formation of Fe₂O₃ concretions and Fe₂O₃-cemented sand grains. Between these 2 extremes there may exist a great variety of pans of varying degrees of intensity. A study of the behavior of Fe(OH)₃ and Al(OH)₃ in the presence of SO₄ and Cl ions indicates the possibility of a sol. state beyond the point of complete coagulation. M. S. ANDERSON

The potash requirement of South African soils. I. DE V. MALHEIRE AND M. H. SLABBER. *South African J. Sci.* 27, 236-52(1930).—Analytical data are given for a large no. of South African soils. The detns. made include mech. analyses, pH , K₂O sol. in strong HCl and in citric acid, K₂O by the Neubauer seedling method, and Al₂O₃ sol. in HCl. The soils are, for the most part deficient in K₂O. Fertilization with K₂O prevents leaf scorch on certain varieties of fruit trees and berry fruits under and conditions and increases the quality of the fruit and its keeping qualities. Fairly const. ratios exist between K₂O dissolved in concd. HCl, citrate-sol. K₂O, and K₂O by the Neubauer method. Approx. crit. values of these several ratios are proposed for these soils when grain crops are grown. The sandy soils are particularly deficient in K₂O, and frequent applications involving relatively large quantities of K₂O are required for

markedly increasing the quality of fruit and for preventing leaf scorch under the arid conditions prevailing

M. S. ANDERSON

The nature of soil acidity as affected by the SiO_2 -sesquioxide ratio. L. D. BAYZA AND G. D. SCARFITT. *Soil Science* 31, 159-73 (1931).—A study is made of the colloidal material from 21 different soils representing various kinds and states of weathering. The SiO_2 -sesquioxide ratio, the total base-exchange capacity, and certain characteristics of the nature of the soil acids have been detd. on each of these colloids. The nature of the soil acids varies considerably in different soils. In weathered soils the nature of the acids is solely a function of the kind and extent of weathering and is independent of the parent material. This indicates that there is more than 1 type of soil acid. Colloidal material of high SiO_2 -sesquioxide ratio extd. from well weathered soils tends to be more highly buffered and to exhibit stronger acidity than colloids having a low ratio. Total exchange capacity is also higher. The nature of the soil acids may prove a valuable criterion in the classification of soils. Buffer capacity of the colloids appears to be primarily a function of the nature of the soil acid. The exchange complex may develop by removal of certain constituents from the original aluminosilicate minerals, by mutual flocculation of colloidal oxides of Al, Fe and Si, and by pptn. of Al, Fe and Si from soils. There is no one direct relationship between SiO_2 -sesquioxide ratio and the total exchange capacity of the colloid, or the nature of the soil acid. Free oxides of Al, Fe and Si may be present in colloids, and it is suggested that it might be desirable to eliminate these when considering the relation of the compn. of colloidal material to physicochem. properties.

M. S. ANDERSON

The development of roots related to the calcium content of soils. JOH. GÖRNING. *Zuckerzärbau* 12, 125 (1930), *Listy Cukrovár* 49, No. 12, Rozhledy 9, Cl. C. A. 25, 353.—The growth of fine roots was studied in soils of various pH values, the greens were normal in soils in which the roots were stunted. In the more acid strata, the roots became stunted regardless of the depth of stratum in which they were planted. In neutral or alk. strata, a luxurious development of roots occurred.

FRANK MARESH

The influence of drying the soil on the microbiological processes in the soil. V. K. ZALESSEKI AND A. M. KULHARKOVA. *Pochvovedenie* (Pedology) 23, Nos. 3-4, 84-111 (in German 111-2) (1928).—Drying the soil increased the soly. of org. matter, the amt. of available N and P compds. Peptone added to dried soil upon wetting did not increase the activity of the flora, whereas mannitol and glucose did. Drying the soil decreases the absorption power of the soil for microbes, hence there is a greater activity providing available org. matter is furnished. The addn. of carbohydrates to a dried soil stimulates the activity of the microorganisms, and they use the available N and P. The disappearance of nitrates from the soil is due primarily to the absorption by plants and microorganisms.

J. S. JORRS

The fungus flora of the soil. H. L. JENSEN. *Soil Science* 31, 123-58 (1931), Cl. C. A. 25, 1934.—A study of the fungus flora of 100 Danish soils of reactions varying from pH 3.34 to 8.35 shows the largest amt. of mycelium in acid soils rich in org. matter. There is no clear relationship between no. of fungi and soil type or soil reaction except that very heavy clay soils are relatively low in fungi. The addn. of lime does not markedly depress fungi, but greatly stimulates bacteria and actinomycetes. Some fungi will grow in soil made as acid as pH 1.5, while others are checked at pH 3.7 to 4.2.

M. S. ANDERSON

Soil bacteria. II. YOSHIEHIKO YAMAMOTO. *J. Agr. Chem. Soc. Japan* 6, 873-83 (1930).—Mucors are classified into 2 groups: (1) monomucors (sporangiophores not branched) and (2) polymucors (sporangiophores branched). The list of soil Mucorales was mentioned.

Y. KIHARA

Reducing the consumption of sulfuric acid in reworking of phosphorite of Izyum. M. A. EGOROV. *Ukrainskii Khim. Zhur.* 5, No. 2, Tech. Pt. 43-63 (1930).—Agricultural expts. have shown that a certain part of Ca contained in the phosphorite of Izyum (I) possesses unusual mobility and acts as a very strong retrograding agent on the phosphoric acid (II) in I, rendering it sparingly sol. and thus only slightly available for plant vegetation. Instead of the costly process of removing Ca from I (Rozhdestvensku, C. A. 22, 2022) the present investigation gives methods of chemically binding Ca, whereby its deleterious action on the mobility of II in I may be entirely eliminated. Conclusions. The activation of II in I by citric acid (III) is greatly increased with the addn. of oxalic acid (IV), KCl and NaCl. The activation of II by IV is increased with NaCl. Under some conditions IV *per se* can entirely remove II from I. The activity of H_2SO_4 on I is greatly increased by addn. of alkali oxalates (which is not the case with phosphorites of other origins), whereby the entire II in I may be rendered H_2O -sol. and the consumption of H_2SO_4 may be reduced about 2.5 times. With the aid of a

little H_2SO_4 and some $(COONa)_2$; a phosphate was prepd. which produced in 1929, with cultivation of sugar beets, results equal to those of superphosphate, the prepn. of which consumes 3 times as much H_2SO_4 . The work is being continued. CHAS. BLANC

The electro dialysis of phosphorite of Izium. K. N. TARANOV. *Ukrainskii Khim. Zhur.* 5, No. 2, Tech. pt. 65-75 (1930), cf. C. A. 23, 460.—The investigation has been undertaken for the reworking of low grade phosphorite (1) of Izium by electro dialysis, whereby its P_2O_5 may become available for plant vegetation. Conclusions: By the electro dialysis method it is possible completely to decompose 1 and sep. all of the P_2O_5 and Ca in the form of aq. solns. The residue of 1 has an acid reaction ($AcONa$). By the double-chamber process of sepn. of P_2O_5 from 1, there are formed residua of 1, which proved as effective as superphosphates in vegetative expts. (oats). The enrichment of residua with P_2O_5 is due to the removal of some components, such as Ca, CO_2 , Cl, etc. By the electro dialysis method only H_2O and the elec. current are required. Questions pertaining to such factors as tension and strength of the current, the relation between the electrode surface and the mass of 1, etc., have not yet been studied.

CHAS. BLANC

Experiments on the production of thermophosphates from the Khibinsk apatite. S. I. VOLPKOVICH AND S. S. PERELMAN. *Udobreniia i Urozhai (Fertilizers and Crops)* 2, 570-80 (1930).—By a flotation method a product was obtained from the Khibinsk apatite with a 40.5% P_2O_5 content. This and 2 other samples contg. 27.12 and 19.47% P_2O_5 were fused with Na_2CO_3 . By using 30.2 parts of the 40.5% P_2O_5 sample for 100 parts of the phosphate at 900-1200° a product was obtained with a 70% coeff. of decompn., the latter representing the ratio of citrate sol P_2O_5 to the total. With a 25% excess of Na_2CO_3 the coeff. increases irrespective of the temp. from 900° to 1200°. With a 50% increase the proportionality between the Na_2CO_3 and increase in P_2O_5 disappears. At 1100° the coeff. drops from 75.3 to 67% and then again increases at 1200° to 82.6%. At 1300° the mixt. fuses but there is no increase in P_2O_5 . The optimum conditions were found to be at 1200° with 1.5 times the amt. of soda used originally, i. e., 45 parts of Na_2CO_3 to 100 parts of apatite. With the 27.1% P_2O_5 sample at 1000-1200° with 20 parts of Na_2CO_3 to 100 parts of phosphate the coeff. of decompn. was 70-80%. A 25% increase of Na_2CO_3 increased the coeff. of decompn. almost to 100%. Even at 900° such a mixt. gave an 82.2% coeff. With the 19.5% P_2O_5 sample at 1200° with 22 parts of Na_2CO_3 to 100 parts of the phosphate almost 100% of the phosphate became citrate sol. Addn. of 10% SiO_2 to the mixt. slightly increased the coeff. of decompn. SiO_2 alone had no effect. An increase of the SiO_2 to 20% produced neg. results. Addn. of $CaCO_3$ alone had no effect, but with the Na_2CO_3 an increase in the coeff. took place. The time factor of heating the mixt. brought out the fact that 30 min. gives the optimum. The cooling of the mixt. immediately after 30 min. heating increased the soly. of the product. Heating with various chlorides gave neg. results. With Na_2SO_4 and charcoal a product was obtained with a coeff. of 80%. Vegetation expts. with thermophosphates show that it is just as good as Thomas slag, but it is slightly inferior to superphosphate.

J. S. JOFFE

The relation between the geological origin and the available phosphoric acid of a soil. L. NEUBERGER. *Kunstdünger u. Leim* 28, 77-8 (1931).—Review. C. J. S.

A colorimetric determination of phosphoric acid in soils. A. NĚMEC, LANÍK AND KORROVÁ. *Vestník státního výzkumného ústavu zemědělského* 5, 18 (1930); *Lísty Cukrovár* 49, No. 4, Rozhledy 2.—The König-Lemmermann method requires a long time for analysis, and soils high in Ca make the ignition of citrate salts very difficult. For routine work the method is: shake 10 g. of soil with 100 cc. of a 1% citric acid soln. (soils high in Ca have to have a correction factor) for 1 hr. Repeat the process for 1 hr. on a subsequent day. Filter or centrifuge, treat 5 cc. of the clear soln. with 10 cc. of 40% H_2SO_4 , bring to a boil, and treat with 0.1 N $KMnO_4$ until a faint but permanent pink color remains. Remove any excess of $KMnO_4$ with H_2O_2 . Digest the soln. on a water bath for 30 min., drain into a 100-cc. vol. flask, treat with 4 drops of α -dinitrophenol, and neutralize with 15% NH_4OH . Cool the soln. and treat with 2 cc. molybdic acid and 0.3 cc. $SnCl_4$. This produces a blue color, make up the soln. to 100 cc. and compare colorimetrically. The standard soln. is made from K_4HPO_4 contg. 0.05-0.5 P_2O_5 per 10 cc. The values run higher than those with the K.-L. method, N. ascribes this to a more thorough extrn.

FRANK MARESH

The use of the Neubauer method for determining phosphoric acid supplies in soils. A. NĚMEC. *Zemědělský arch* No. 3, 129 (1930), *Lísty Cukrovár* 48, No. 52, Rozhledy 11.—The Neubauer method of sprouting plants yielded low P_2O_5 values in soils which were adequately supplied according to fertilizer tests. The cause of a lowered resorption of P was found to be due to a low N content. All soils had a low N and low nitrifica-

tion power. The Neubauer method does not take into consideration the laws of minima. Fertilizing with N increased the resorption of P in which low Neubauer values were obtained. FRANK MARSH

Efficiencies of phosphoric acid of various fertilizers for spring- and autumn-sown barley. IWAO OMAI *Ann Agr. Expt Sta Government-General of Chosen* 3, 394-402 (1930) — Application of NaH_2PO_4 , bone meal and superphosphate gave about 190% greater yield and 130-140% more of absorbed P_2O_5 ; rice polish and AlPO_4 gave 150% greater yield and 60-70% more absorbed P_2O_5 than when no P_2O_5 was applied. In general autumn sown barley showed higher absorption and higher yield than spring-sown barley. Tests for residual P_2O_5 after the first crop showed great value for the second crop. K. KRISOTA

A contribution to the evaluation of requirements for fertilizing soils with phosphorus. A. NEMEC *Vestník Ústř. Jednoty Repato Čs* 11, 277 (1930), *Listy Cukrovár* 49, No. 4, Rozhledy 1 — The use of P_2O_5 in fertilization is closely allied with Ca; soils with adequate Ca should have more available P_2O_5 than soils deficient in Ca. Other elements controlling the availability of P_2O_5 are Fe and Al, soils contg. more than 50 mg Fe freely sol. in 1% citric acid fail to respond to P_2O_5 treatments, even though they show a P_2O_5 deficiency by the König Lemmermann method. The method for detg. the available P by evaluating only the P_2O_5 sol. in H_2O becomes of no value for soils contg. less than 3 mg P_2O_5 (sol.) per 1 kg soil. Soils contg. large quantities of Fe did not react to P_2O_5 fertilizations, even though the sol. P_2O_5 was high or low. All available methods were inadequate to det. what soils would respond to P_2O_5 treatments. Soils deficient in both Fe and P_2O_5 gave increasing yields of beets, N attributes this to SiO_2 , for only soils contg. less than 12 mg sol. SiO_2 per 1 kg of soil reacted to H_3PO_4 treatments regardless of whether the P_2O_5 was high or low. In soils with a high Fe content, P_2O_5 fertilizers fail to increase the yield of beets, even though the sol. SiO_2 may be less than 12 mg per kg of soil. FRANK MARSH

The retrograding action on phosphates of the soil under various conditions. P. V. TRIMPOV *Ukrainskii Khim. Zhur* 5, No. 2, Tech. pt. 77-91 (1930) — The influence of moisture on the retrogradation of phosphates is considerable. Thus the soly. in H_2O of CaH_2PO_4 decreases by 30% with an increase of 30-100% in the moisture based on the total moisture capacity of the soil. The soly. of Izyum phosphate in citric acid increases in 65% with an increase in moisture. The highest amt. in soln. was observed at a moisture content of 80-100% of the total moisture capacity. The influence of temp. on the retrogradation is different for different phosphates, e.g., temp. changes do not affect solns. of Izyum phosphoric acid in H_2O or in citric acid, while increase in temp. greatly decreases the soly. of CaH_2PO_4 in both water and citric acid. The retrograding action on phosphates of soils of the region of forest steppe varies greatly, there is a regular decrease of retrogradation with declining richness of the soil. CHAS. BLANC

The constitution and citrate solubility of tricalcium phosphate and of phosphate rock. K. D. JACOB *Phosphorus Digest* (April, 1931), pp. 7-9; cf. *C. A.* 25, 1619 — When 0.5-g. samples are used, 65-70% of the total P_2O_5 in pptd. $\text{Ca}_3(\text{PO}_4)_2$ and in steamed bone meal is sol. in neutral N.H. citrate soln., but when the wt. of the sample is increased to 2.0 g., only 25-30% of the total P_2O_5 goes into soln. The citrate soly. of ground phosphate rock is only about 12% of that of $\text{Ca}_3(\text{PO}_4)_2$ and bone. The low soly. of phosphate rock seems to be due principally to the fact that the phosphate is combined with F to form a complex Ca fluorophosphate of the same type as cryst. fluorapatite. K. D. JACOB

Can superphosphate and its components increase the solubility of soil potash? C. DREYSPRING AND K. BRAUNEMANN *Superphosphate* 4, 69-75 (1931), cf. *Dreyspring and Henrich, C. A.* 25, 1935 — The increased soly. of soil potash as a result of the application of superphosphate is due principally to the presence of free H_3PO_4 in the superphosphate. The $\text{Ca}(\text{H}_2\text{PO}_4)_2$ and CaSO_4 in superphosphate also increase the soly. of soil potash. K. D. JACOB

Superphosphates enriched with ammonia. S. I. VOLPKOVICH, L. E. BERLIN, I. L. HOPMAN AND A. A. IOVAS *Udobrenie i Urozhai (Fertilizers and Crops)* 2, 435-504 (1930) — Tests were made of mixing H_2SO_4 with $\text{NH}_4\text{H}_2\text{PO}_4$, with $(\text{NH}_4)_2\text{HPO}_4$ with a mixt. of both, and with $(\text{NH}_4)_3\text{PO}_4$. The factors studied were the quant. relationship of the reagents, order of mixing, duration of process and temp. involved. A definite amt. of H_2SO_4 was placed in a Pb container, and the various NH_4 phosphates were added gradually. A stirrer was started and raw phosphates were added slowly, lumping being avoided. The stirring lasted 1-2 min., after which the mixt. was placed in a thermostat at 95-105°. After 3 hrs. the mixt. was weighed, to follow the loss in wt., and placed in thin layers on paper, it was left that way for 2 days and weighed again. Analy-

ses were then made. Total H_2O -sol and citrate-sol P_2O_5 were then detd. Methods of prep. the various NH_4 phosphates, as well as the amts of H_2SO_4 used with them, are given. J S JOFFE

Organic fertilizers. VII. KIYONAO YOSHIMURA, KOTARO NISHIDA AND ARITOMO YAMADA. *J Agr. Chem. Soc., Japan* 6, 999-1002(1930).—Steamed bone powder was kept with water for 55 days. The av temp was 32.8° . The aq ext of the putrefaction products was analyzed: total N 2.550%, protein N 0.044%, P_2O_5 0.019%. Ninety % of total N was changed into NH_4 . One g histamine (Au salt), 3.31 g putrescine-HCl and 280.74 g NH_4OH were isolated from 10 kg of the bone powder. Y K.

The economical and profitable use of commercial fertilizers under present conditions. O ENGELS. *Kunstdünger u. Leim* 28, 126-36(1931).—A general discussion. C J SCHOLLENBERGER

The use of sawdust as a fertilizer. MARCEL H MOTTE. *J agr. prat* 95, 192 (1931).—Sawdust used exclusively in large quantities on the soil has generally given an excellent plant growth without any apparent lack of nitrate and with normal nitrification. Artichokes, leeks, sugar beets, maize and strawberries all have produced good crops without any other fertilizer. J R ADAMS

The Gerard manure pit. F MAGNIEN. *J agr prat* 95, 71-4(1931).—The conditions governing the transformation of barnyard manure into a fertilizer contg available nutrients are listed. The Gerard manure pit is intended to satisfy these conditions and is described in considerable detail. J R ADAMS

Potash fertilizer for sugar beets. J-PH WAGNER. *Compt rend acad agr France* 17, 77-9(1931).—Sugar beet cultivation calls for a soil with a pH of 7-7.5, which can be obtained by liming. Max yields have been obtained when approx 1,000 kg/acre of sylvinite lamite has been added to the soil. The yields have amounted to 761 kg of sugar per acre. The large amts of K_2O aid the plant in obtaining an early and vigorous growth and greater resistance to disease. They also sensibly increase the saccharin content. J R ADAMS

The nitrification of ammonium sulfate as influenced by soil reaction and degree of base saturation. JAMES A. NAFTEL. *J Am Soc. Agron.* 23, 175-85(1931).—There appears to be a relationship between the percentage base and Ca satn and the nitrification of $(NH_4)_2SO_4$. After 10 days' incubation nitrification occurred at pH 4.1 only in the soil with a high percentage base and Ca satn. Lower base satn prevented nitrification. All the soils had a higher percentage nitrification of $(NH_4)_2SO_4$ at the highest pH values studied. In most cases plant growth was better at pH 6.5 than at the very acid reactions. Nitrification varied considerably in different soils and showed a fairly good correlation with plant growth data. J R. ADAMS

Experiments with calcium nitrate-I. G. F. X. KREŠL. *Listy Cukrovar* 49, 261-2(1931).—The $Ca(NO_3)_2$ was prepd by the I. G. firm and contained N 15.5% and CaO 28%, which corresponds to 50% $CaCO_3$, the N is rapidly utilized by plants. Expts were carried out in 18 sections of an estate, with varied results. Because of the sol form of N, less H_2O is required for soln than for any other nitrate fertilizer, the $Ca(NO_3)_2$ maintains a porosity during periods of drought, and on account of its fine dispersion the Ca is equally distributed and swells and aerates the soil. Plots fertilized with 400 kg superphosphate and 200 kg of a 40% K salt per ha. showed an increased yield of sugar beets over unfertilized plots, adjacent plots fertilized with 100 kg $Ca(NO_3)_2$ per ha. showed about the same increase as those with the superphosphate mixt. F. M.

The action of chromium on the growth of plants. EL. HASELHOFF, F. HAUN AND W. ELBERT. *Landw Vers Sta* 110, 283-6(1930).—The addn of Cr to the soil has a detrimental action on the plant growth as shown by the effect on the yield of summer barley and mustard. The amts. used were not greater than 0.003%. This action depends on the compd of Cr used and is greater with chromic acid than with chromic hydroxide. The injurious action is greater in sandy than in loam soils. J R H.

The action of arsenic on plant growth. E. HASELHOFF, F. HAUN AND W. ELBERT. *Landw Vers Sta*, 110, 287-9(1930).—Addns of As up to 0.003% have no effect on the yield of summer barley. This result is not in disagreement with the work of König because different compds and amts of As were used. JOHN R. HILL

The nutrition requirements of vegetable plants. I. Cabbage, beet, carrot and onion. FELIX KOTOWSKI. *Roczniki Nauk Rolniczych i Lesnych* 24, 372-449 (450-2 English) (1930).—Fertilizer expts. carried out on a rather poor sandy loam in 1928 near Warsaw are described. Beets, carrots and onions were sown in exptl plots in rows 40 cm apart. The food supply consisted of mineral fertilizers and stable manure in the ratio of 600 kg per 100 sq m after the tomato crop. Cabbages and carrots received 90 kg. N, 40 kg. P_2O_5 and 120 kg K_2O , beets 63 kg N, 28 kg P_2O_5 and 84 kg K_2O , onions 28

kg P_2O_5 84 kg K_2O and 63 kg N per hectare. N was applied as NH_4NO_3 , P as superphosphate and K in the form of 20% potash salts. Sampling for chem. analysis was done at various intervals of plant development, the last sample being taken at harvest time. The results are presented in tables and the following conclusions are drawn. For (a) Brunswick short stem cabbage, (b) dark red flat Egyptian beet, (c) Chanteney carrot and (d) Zittau onion the max. intake of N, P, K and Ca falls in a in the second b in the fourth c in the third and d in the fourth months after transplanting of (a) and germination of (b, c, d). A crop of (a) 500 000 kg heads and 370 000 kg refuse parts (b) 500 000 kg roots and 200 000 kg leaves, (c) 250 000 kg roots and 80 000 kg leaves (d) 294 000 kg bulbs and 72 000 kg leaves per hectare removes from the soil 150 122 80 80 kg N, 31 31 30 26 kg P_2O_5 , 225 200 120 116 kg K_2O , and 180 76 100 58 kg CaO resp. The N P_2O_5 and K_2O ratios are 3 1 4, 3 5 1 5 5, 2 5 1 4 and 3 1 4 5 resp. As compared with cereals, calingeta, beets, carrots and onions are remarkable for their high N, P, K and especially Ca contents. There is also a considerable difference in the N P_2O_5 and K_2O ratio, which is on an av. 2 1 2 in cereals and 3 1 4 in the plants reported here. These differences, therefore, should be considered in any fertilizer program. J KLČERA

The role of humic materials in fertility. A DEMOLON. *J. agr. prat.* 95, 109-11 (1931).—The maintenance of humic reserves in the soil is almost always indispensable in maintaining a high fertility and in particular for obtaining the max. use from mineral fertilizers. They are excellent soil conditioners and aid the soil in retaining H_2O . Colloidal humic material increases the nutritive absorptive capacity of soils. The value of farm manure and the manuf. of artificial manure are discussed. J R ADAMS

Brief study of (reasons for) beneficial action of brown coal upon the development of cultivated plants. II. ALFRED KISSEL. *Brennstoff-Chem.* 12, 101-7 (1931), cf. C. A. 25, 371.—Investigations during 1924-25 with grain growing in arable soil fertilized with coal dust show a growth increase by kernel weight of 24.55%, depending upon the amt. of coal dust used. This is believed due to improved germination and branching rather than to any change in soil compn. Potatoes also showed tuber-weight increases but with sugar beets the sugar content fell off with root-weight increases. Structural aggregates in the soil are formed by the freshly pptd. water-insol. humic acid salts which lose this cementation property on drying. The impregnation of soil with coal dust or humic compds. thus increases the resistance to water influx into individual soil particles and increases water seepage and aeration in upper soil layers and, in general, restores its fallowness. Small quantities of coal dust increase the soil resistance to machinery but large amts. decrease it. It is contended that difficulties encountered in the use of artificial agricultural fertilizers are to be traced to factors other than the amt. of mineral matter made available. The structural compn. of the soil, its phys. characteristics and the total content of org. humic combinations are also involved. F W JUNG

Decomposition of rotenone in solution. HOWARD A. JONES. *Ind. Eng. Chem.* 23, 387-8 (1931).—The decompn. of rotenone is accelerated by access of air but is not affected by differences in amt. of light or by the presence of water. Pyridine solns. of rotenone decompose most rapidly and in a few days crystals of dehydrorotenone and rotenonone are formed by oxidation. Rotenone is completely sol. in benzene, and in a benzene soln. there is no appreciable decompn. Acetone may be used for a water-sol. solvent. J R ADAMS

The relation of chemical composition of cultivated and wild green-manure plants to decomposition of the nitrogenous constituents. HIROO MISU, IWAO OMAT and TADAO HIBINO. *Ann. Agr. Expt. Sta. Government-General of Chosen* 5, 1-126 (1930).—Twenty-two varieties of cultivated and wild green-manure plants were decomposed by fermentation for 48 days under conditions simulating those in soil with upland and lowland systems of culture. The relations of ammonification, nitrification and the yield of titratable acids to the chem. compn. of green manures are as follows. Ammonification differs greatly with the different raw materials used under both upland and lowland conditions. In upland conditions, it decreases later. Nitrification is negligible under lowland conditions, but it differs greatly under upland conditions with the different raw materials used, and it decreases later. Attainment of max. ammonification is favored by the increase of total N, ash, CaO, or CaO + MgO and disturbed by the increase of crude fiber, N free ext., crude-fiber protein, and N free ext./protein. Attainment of max. nitrification is favored by the increase of total N, K_2O and CaO + MgO and disturbed by the increase of crude fiber, N free ext., and N free ext./protein. Attainment of max. titratable acidity is favored by the increase of total N and of CaO + MgO and disturbed by the increase of lignin, N free ext., and N free ext./protein. The max. ammonification and the

development of max acidity were greater in lowland conditions than in upland conditions

K. KITSUTA

Availability of nitrogen of green manure for rice and the supplementary value of various fertilizers. HIDEO MITSU AND H. SUMIYAMA. *Ann Agr Expt Sta Government General of Chosen* 4, 65-94(1929) —The availability of N was lower in dried than in fresh green manure. The value of N of dried green manure was greatly increased by replacing $\frac{1}{2}$ of the N with $(\text{NH}_4)_2\text{SO}_4$. Although the N of fresh material appeared less available, the yield obtained was better than that with fermented green manure. Superphosphate, K_2O and CaO had great value as supplements to green manure, in increasing availability of the N and yields. Yields with fresh green manure were 83.4 and 89.3% of those with soy bean cake and $(\text{NH}_4)_2\text{SO}_4$ resp.

K. KITSUTA

The availability of nitrogen of air-dried green manures for rice. HIDEO MITSU. *Ann Agr Expt Sta Government General of Chosen* 4, 167-72(1929) —Green manures showed lower availability of N and lower yield than did soy bean cake and $(\text{NH}_4)_2\text{SO}_4$.

K. KITSUTA

Nitrogen in the intensive exploitation of pastures. CAMILLE MATIGNON. *Chimie et Industrie* 25, 271-5(1931) —A discussion is given of the advantages, from the double standpoint of yield and aqua lity, of the use of N fertilizers on pasture land. A. P. C.

Comparative value of nitrogenous fertilizers on the growth of fall and spring barley. IWAO OMAI. *Ann Agr Expt Sta Government General of Chosen* 3, 317-54(1930) —Among 15 tested materials Chukan NaNO_3 and $(\text{NH}_4)_2\text{SO}_4$ gave the highest percentage of recovery of N by the plants (60-66%) and the best yield. CaCN_2 , dried fish meal, dried blood, soy bean cake, bone meal and human excrement showed 45-55% recovery of applied N by the plants with 68-87% yield compared with the yield by $(\text{NH}_4)_2\text{SO}_4$ as 100%. Heavy application of mulch made of weeds was rather injurious on the first crop as a result of its fermentation, and the benefit to the second crop was small in comparison with the loss of N by fermentation. The efficiency of the N in stable manures from swine and cattle were very small. The value for the second crop of any residual N after the first crop was rather small.

K. KITSUTA

The effect of various fertilizer treatments on the crop-yield factors and the structure of leaves of barley and wheat. K. HORKNOLT. *Z Pflanzenernähr, Düngung u Bodenk* 9B, 289-300(1930) —Stooling, weight per 1000 grains, grains per single head, weight of grains per head and total yield are affected by various single fertilizers and combinations. K has little effect on stooling. P more and N most. The 1000 grain weight and weight per single head are least with N. A mixt. of N, P and K usually gave the highest values for all the factors mentioned and always the highest yield. The leaf surface varies with the fertilizer treatment, e. g., if no treatment is 100, then a mixt. of K and P is 97.8, a mixt. of K and N 128.9, a mixt. of P and N 144.4, a mixt. of K, P and N 134.7, and 2 pts K, 1 pt P and 1 pt N 130.7. The microanatomic structure also varies with different fertilizer treatments.

H. B. SHU-SH

Effect of ammonium sulfate on field germination of rape and turnips. J. W. WOODCOCK. *New Zealand J Agr* 42, 89-90(1931) —Germination of rape and turnip seed fertilized with $(\text{NH}_4)_2\text{SO}_4$ either alone or mixed with other fertilizer materials, at the rate of 100 lb/acre was only 10 to 44% of the germination obtained when the seeds were fertilized only with a mixt. of 100 lb of superphosphate plus 100 lb of ground limestone. Fertilization with superphosphate alone gave germination counts 66 to 100% of those obtained with the mixt. of superphosphate and ground limestone. In all cases the fertilizers were applied with the seed in 14 in. rows.

K. D. JACOB

Fertilizer experiments on citrus seedlings. I. TAKAHASHI. *J Okitsu Hort. Soc (Japan)* 25, 38-50(1930) —Pot expts with seedlings indicated that N is the most necessary element for growth. The requirements of P_2O_5 and K_2O were $\frac{1}{2}$ and $\frac{3}{4}$ resp. of the N requirement. Requirement of P_2O_5 for growth of the citrus tree was not very great, but deficiency of this element retarded growth markedly. The requirement of K_2O was greater than that of P_2O_5 , however, the growth continued uninterruptedly without K_2O for a certain time by compensation with the supplement of the other elements.

K. KITSUTA

Fertilizer experiments with Satsuma orange. K. TASEKI, Y. INOUE AND N. MATSUOKA. *Imp Hort Expt Sta (Japan), Research Bull* 13, 60 pp (1930) —Plots were arranged into those with no fertilizer, N, P or K and complete fertilizer treatment. No fruit was obtained from the trees in the plot without N. When P was omitted, the yield was much decreased, and the acidity of the fruit was high. Lack of K caused decrease in yield, poor color of skin and small sized fruits. The spoilage of fruits during storage was 5.5 times higher than that of the fruits from the completely fertilized plot.

K. KITSUTA

Is fertilization to prevent frost injury practical? WALTER OBST. *Kewitzinger u. Linn* 28, 29-7(1931)—The resistance of plants in general toward injury from cold is increased by complete fertilization, excess N being avoided. C. J. S.

The importance of artificial fertilizers in modern pond culture. K. HEITZ. *Kewitzinger u. Linn* 28, 27-8(1931)—Growth of plankton serving as food for fish is greatly increased by proper fertilization with K_2O and P_2O_5 . The application of CaO to a drained pond during winter is advised as a sanitary measure. C. J. SCHOLLEVERGAS.

Carbon dioxide in relation to glasshouse crops. V. An analysis of the response of the tomato crop to an atmosphere enriched with carbon dioxide. H. L. WATTS. *Ann. Appl. Biol.* 17, 755-61(1930)—The enrichment of the atm. has the following effects upon the tomato plant. It shortens the period between opening of the flower and picking of the fruit. A high percentage of blossoms develop into fruit. Less fruit is retarded in ripening. There is better development of the truss. There is earlier development and ripening of trusses. There is better swelling of fruit at end of season. Conclusion. Increased CO_2 concn. is most beneficial to the plant at periods of low fruit production. C. H. RICHARDSON.

A study by foliar diagnosis of the influence of temperature on plant nutrition. H. LAGAYE AND L. MAUME. *Compt. rend.* 190, 1516-8(1931); cf. *C. A.* 24, 4575; 25, 2534—Potatoes were grown on plots receiving various fertilizer treatments, and the leaves were analyzed for P at intervals. During May the temp. averaged about 14° , and the P content remained rather const. During June, when the temp. averaged 21° , the P content decreased markedly on the plots which received N, but did not on those without N. The explanation offered is that with added N tuber development is more abundant, and this tends to deplete the P in the leaves. J. J. WILLAMAN.

Some studies on the use and action of mercuric chloride as a fungicide. J. L. HOWARD. *Twenty-first Ann. Rept. Quebec Soc. Protect. Plants* 1928-9, 40-2(1929). *Rev. Applied Microb.* 10, 195-7—The control of fungicides in certain seed borne diseases by use of $HgCl_2$ is due to retardation of the germination of the spores sufficient for them to be outgrown by the host and so removed from the meristematic tissue. Stimulation of the seed after treatment is attributed to indirect partial sterilization effect on the soil. Temp., concn., time of action, age of spore and host plants, assocn. and type of seed are considered to affect the toxicity. OWEN E. STEPHARD.

Titrimetric determination of arsenic in plant-protecting agents. J. PÁSKY. *Magyar Chem. Folyóirat* 35, 76-83, 95-101(1930)—The agents are destroyed with HNO_3 and H_2SO_4 , then added to a concn. of about 50% H_2SO_4 . Now 50 cc. concd. HCl , 5 g. crys. FeO_4 , and 2 g. KBr are added, 25 cc. is added, and As titrated in the dilute according to Gayer with 0.1 N $KBrO_3$ soln. Detns. may thus be made within 1-5 hrs. Also a micro method was worked out, the results of which quite agree with those of the macro method. Hydrazine sulfate should replace the FeO_4 in the macro-dn. S. S. DE FINALLY.

The biochemical influence of arsenic. J. PÁSKY. *Magyar Chem. Folyóirat* 36, 111-8(1930)—Spores of *Tullesia tricus*, Berk. Winter, adsorb some As of H_2AsO_4 and arphenamine but no As of other arsenicals, either inorg. (arsenites, arsenates) or org. (arsacetin, stoxyl, solarson, cacodylic acid, etc.). Adsorbed As does not kill spores but hinders their multiplication. Agents of which no As is adsorbed have no influence at all on spores. Spores treated with arsenates and dried without washing off the agent were killed. Arsenates did not show such results. S. S. DE FINALLY.

Effect of arsenic on the composition of citrus fruits. I. TAKAHASHI. *J. Ohitsu Hwt. Soc. (Japan)* 25, 153-63(1930)—Spraying citrus trees with arsenates caused a decrease in acidity and an increase in sugar content of fruit juice. The reduction of acidity depends greatly upon the time of spraying, amts. of other spraying materials mixed with arsenate and the method of application. Excessive use of arsenate should be avoided, as it affects the yield and taste of fruit. A decrease in acidity and an increase in sugar content of fruit caused poor keeping quality, therefore, in warmer regions caution must be exercised in spraying with arsenicals. Beside controlling harmful insects, spraying arsenates may be of benefit by hastening maturity or reducing excessive acidity of fruit. K. KITSUTA.

Effects of some sulfur sprays on fruit-bud development. C. E. FETCH AND J. L. HOWARD. *Twenty-first Ann. Rept. Quebec Soc. Protect. Plants* 1928-9, 39(1929); *Rev. Applied Microb.* 10, 197.—Comparative tests of koloform (prepd. by fusing bentonite with S, with Ca caseinate added to assist the mixing in water) and lime-sulfur on the development of apple trees shows that koloform treated buds were 0.4903 in. long when those sprayed with lime-sulfur were 0.3979 in. long. This difference is attributed to absence of injury in koloform treated foliage, whereas burning and dwarfing of foliage

was evident on the trees treated with lime sulfur. Koloform was applied at the rate of 8 lbs., and lime-sulfur at the rate of 1 gal. to 40 gals. water. OGDEN E. SHEPPARD.

The mineral nutrition of plants in relation to their predisposition or resistance to attack by pathogenic agents. L. PETRI. *Boll. R. Staz. Pat. Veg. N. S.* 10, 121-52 (1930), *Rev. Appl. Mycol.* 10, 200.—The action of different fertilizers (including nitrogenous and phosphatic fertilizers, lime, sulfur, gypsum, potash, iron and aluminum sulfates, and manganese salts) on the anatomical and physiol. properties of plants with reference to their resistance to injurious inorg. agencies as well as to various fungus and animal parasites was studied. In general, nitrogenous dressings aggravate susceptibility, whereas phosphatic and potash dressings reduce it. The action of the fertilizer is indirectly on the mechanism of the host resistance and is influenced by local factors, comparative tests must be carried out under identical conditions. OGDEN E. SHEPPARD.

Spraying and dusting experiments with potatoes on Long Island. H. C. HUCKETT. N. Y. State Agr. Expt. Sta., *Bull.* No. 592, 1-38 (1931).—Expts. in spraying and dusting Irish Cobbler (I) and Green Mountain (II) potatoes have been conducted during 5 successive seasons, 1926-1930. With the exception of applications containing nicotine, Bordeaux mixt. (4-6%) was used in sprays and $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ and $\text{Ca}(\text{OH})_2$ (20-80) for dusts. For biting insects CaAsO_4 was added to Bordeaux mixt. at the rate of approx. 5 lbs. or to dust mixts. at 10-12 lbs. per acre. Arsenical treatments were usually made 3-4 times each season to combat Colorado beetle and flea beetles. For sucking insects nicotine sulfate (40% soln.) was added to Bordeaux mixt. or a mixt. of dolomite and $\text{Ca}(\text{OH})_2$ (10-50) at the rate of about 1 pint or 3 pints per acre, resp. There were no significant differences in the results from dusting and spraying in the case of I. Also nicotine treatments in spray or dust form did not give rise to additional increases in yield over other spraying and dusting operations. With II Cu treatments in spray form gave on an average slightly superior results to Cu treatments in dust form. Nicotine treatments in dust form were more reliable, and in years when aphids were a limiting factor gave superior results to nicotine in spray form. J. KUČERA.

The effect of gypsum upon the growth and common scab of the potato. J. G. COULSON. *Twenty-first Ann. Rept. Quebec Soc. Protect. Plants* 1928-9, 64-80 (1929), *Rev. Appl. Mycol.* 10, 202.—Increasing applications of gypsum above 500 lbs. per acre in field tests progressively reduced the number of tubers formed and gave an increased amount of scab infection. A lab. test showed that addition of gypsum did not effect the pH of the soil. Addition of gypsum to potato dextrose agar from 0.5 to 50% caused no difference in the rate of growth of *O. scabies*. OGDEN E. SHEPPARD.

Seed treatment for controlling covered smut of barley. R. W. LEUKEL. U. S. Dept. of Agr. Tech. Bull. 207, 1-22 (1930), *Rev. Appl. Mycol.* 10, 173-4.—Satisfactory control of covered smut (*Ustilago hordei*) of barley was obtained by steeping the seed grain for 1 hr. in either 1 in 320 formaldehyde, 0.5% semesan, 0.5% corona 020, 0.5% nupulun, 0.25% germisan, 0.25% tallantin or 0.5% Bayer compd. Expts. with more than 45 dusts indicated that their effectiveness is independent of soil reaction, but a soil moisture content of less than 25% satn. decreases the efficacy of most dusts. Hochst, abavit B, and ceresan are the most effective dusts and give satisfactory control. Use of dusts avoids the seed injury common with steeping treatments, and thus better stands and yields are obtained. OGDEN E. SHEPPARD.

Dry picking for control of smut in oats. W. B. MILLER AND J. A. MORROW. *J. Dept. Agr. Victoria* 29, 86-9 (1931).—Good control was obtained by dusting the seed with Cu carbonate at the rate of 2 oz. per bushel. Treatment of the seed with formalin solns. gave better control of the smut but adversely affected germination of the seed. K. D. JACOB.

Action of low potencies. W. MAAG. *Sddeutsche Apoth.-Ztg.* 71, 143-5 (1931).—The results obtained in a series of expts. on the germination of wheat in the presence of varying low potencies of HCl , H_2SO_4 and CuSO_4 are illustrated, as also those obtained with paper strips vertically suspended and just dipping into varying potencies of Koff and H_2SO_4 . W. O. E.

Control of the webbing spider (*Lampona obscoena*) of citrus trees. S. A. COCK. *J. Dept. Agr. Victoria* 29, 83-5 (1931).—Good control is obtained by spraying the webs with an emulsion of 2 pints of kerosene and 4 ounces of soap in 3 gallons of water to which is added 1 lb. of pyrethrum. The kerosene serves to penetrate the web, while the pyrethrum acts as the killing agent. K. D. JACOB.

The destruction of weeds by chemical procedures. M. BARNUT. *J. agr. prat.* 95, 9-11 (1931).—Comparisons are made of the efficacy of the weed eradicators used for those weeds which are harmful to cereals. A 6-8% soln. of H_2SO_4 gives excellent results, destroying charlock, chenopodium and polygonum. The objection is that it is

very active in deliming the soil. Dehydrated uron sulfate and a nitrocopper compd are very effective against charlock, but both these materials are somewhat more expensive. A special powd. silvite is perfectly satisfactory against some of the weeds such as charlock, but great care must be used, as large quantities of this material are injurious to the crop. J. R. ADAMS

Cucumber disease investigations on Long Island. E. E. CLAYTON. N. Y. State Agr. Expt. Sta., Bull. No. 500, 3/24/1931.—The cucumber diseases, mosaic, wilt and mildew or blight, can be best checked by treating the seeds with org. Hg compds. when the seed is sown early, by using wet and dry sprays with dust mixts. such as $\text{Ca}(\text{AsO}_4)_2 + \text{CaSO}_4 \cdot 2\text{H}_2\text{O}$ (1-15), copper lime, kayso- $\text{Ca}(\text{AsO}_4)_2$ (3-3-50) and by breeding disease-resistant varieties. Spraying does not control mosaic. J. K.

Study of the warble and its eradication. A. GANSSER. *Schweiz. Leder- und Ztg.* 1930, No 34. *J. Intern. Soc. Leather Trades Chem.* 14, 324, cf. C. A. 20, 2223.—The hypocotyle (nicotine salve) treatment previously described has worked well in practice. The treatment seems to have a lasting prophylactic effect. H. B. M.

Recovering of NH_3 in beet-sugar manufacture [for use as fertilizer] (VOSTOKOV, KOPLYARENKO) 23. Detection of F in soils (RUCKENPOFFER) 7. Organic solvents for aiding the removal of spray residue from waxy or oil-covered fruit (ROBINSON) 12. Estimating the solubility of acid Pb arsenate within the alimentary tract of the silk-worm (CAMPELL, LUKENS) 111. Quinquennial review of the mineral production of India for the years 1924 to 1928. Saltpeter (PASCOE) 18. $(\text{NH}_4)_2\text{SO}_4$ (Hung. pat. 101,000) 18.

KEEN, BERNARD A. The Physical Properties of the Soil. New York: Longmans, Green and Co. 381 pp. \$8.

VERMOREL, V. Agenda agricole et viticole, 1931. Paris: Dunod. 384 pp. F 650.

Reagent for treating alkali soils. CHRYCOT, LTD. Hung. 101,149, Jan. 4, 1930. S bacteria and NH_4 salts are added to a mixt. of powd. gypsum and S. *q. q.* a mixt. of 45% CaSO_4 is made of 45% S and 50.5% powd. peat inoculated with S bacteria.

Fertilizers. CHEMISVERFAAREN G. M. B. H. Fr. 37,452, Sept. 21, 1929. Addn. to 682,081 (C. A. 24, 4351). Crude phosphate is attacked with HNO_3 and the Ca is pptd. as CaSO_4 by a mixt. of K_2SO_4 and MgSO_4 , or a corresponding natural compd. The amt. of $\frac{1}{2}$ g. is equiv. to that of P_2O_5 in conformity with the formula MgHPO_4 . The CaSO_4 is pptd., K_2CO_3 is added to the soln. to ppt. MgHPO_4 , and the rest of the P_2O_5 is pptd. as $\frac{1}{2}$ g. $(\text{PO}_4)_3$ by adding more K_2CO_3 . KNO_3 is obtained by evapn.

Fertilizers. ODDA SMELTEVERK AKTIESELSKAB and ERLING JOHNSON. Fr. 27,494, Oct. 7, 1929. Addn. to 682,423 (C. A. 24, 4351). The soln. contg. phosphoric acid obtained by treating natural phosphates with HNO_3 is neutralized or caused to react with substances forming a fertilizer, such as NH_3 , compds. of NH_3 and CO_2 , urea, cyanamide or its salts or natural phosphates, so that on removing the water a dry stable fertilizer is obtained. Other salts of K, NH_4 , etc., may be added in the solid state.

Fertilizers. EDONARD URBAIN. Fr. 693,221, Aug. 29, 1929. A fertilizer contg. P_2O_5 30-34, K_2O 8-11 and N 10-11% is made by using natural phosphate and aluminosilicate rocks in suitable proportions. Lime is introduced in such amt. that a utilizable slag is obtained.

Fertilizers. I. G. FARWENING A-G. Fr. 37,217, July 16, 1929. Addn. to 673,093 (C. A. 24, 2538). Compd. fertilizers are prepd. contg. NH_4NO_3 but not CaCO_3 .

Fertilizers. HERBERT J. KRASE, HARRY C. HETHERINGTON and LOUIS A. PINCK (to Arthur B. Lamb, trustee). U. S. 1,797,003, March 17. See Fr. 678,693 (C. A. 24, 3599).

Fertilizers. ODDA SMELTEVERK AKTIESELSKAB and E. JOHNSON. Brit. 339,340, Oct. 27, 1928. See Fr. 682,423 (C. A. 24, 4351).

Calcium phosphates and fertilizers. ODDA SMELTEVERK AKTIESELSKAB and E. JOHNSON. Brit. 339,340, Dec. 14, 1928. The mother liquor obtained by treating rock phosphate with HNO_3 and sepg. a substantial proportion of $\text{Ca}(\text{NO}_3)_2$ by cooling as described in Brit. 339,340 (preceding abs.) is treated with gaseous or aq. NH_3 , NH_4 carbonate or alkali metal carbonate; pptd. di. Ca phosphate is sepd., and the filtrate is used for the production of products contg. N and phosphoric acid, or the di. Ca phosphate may be treated together with the soln. to obtain solid products. By suitably restricting the quantity of $\text{Ca}(\text{NO}_3)_2$ sepd., relatively pure NH_4NO_3 can be obtained by evapn. the filtrate. Production of tri. Ca phosphate and various fertilizer mixts. also is described.

Fertilizer comprising urea and calcium nitrate. I G FARBERND A-G Brit. 338,644, Sept 16, 1929. A fertilizer contg the double compd. of urea and $\text{Ca}(\text{NO}_3)_2$ is prepd by melting together the anhyd. components preferably in the mol proportions of 4 and 1, resp., and subdividing by spraying or other suitable treatment. Oil may be added to increase the stability of the product in storage. Cf C A 25, 554

Mixed ammonium compounds suitable for use as fertilizer. R. TERN Brit. 338,669, Aug 22, 1929 Gases produced by roasting crude S or spent oxide from gas and coke-mfg plants are mixed with ammoniacal gases produced by heating the gas-water from the scrubbers, and the material is ionized by a high tension current of 70,000–80,000 v., a mixt being produced which comprises NH_4 sulfate, sulfite and imidosulfonate and other compds. of S and N11₂ which may be elec. pptd

Stimulant for tap roots and tubers. I G FARBERND A-G (Georg Wesenberg and Eduard Buschkopff, inventors) Ger 515,884, Mar 28, 1926 Salts of polythionic acids act as a stimulant and increase potato, turnip, etc., yields Examples state that mixts of $\text{Na}_2\text{S}_2\text{O}_4$ or BaS_2O_4 or CuS_2O_4 , $\text{Na}_2\text{S}_2\text{O}_4$, with kieselguhr, cause 20–50% increases in potato crops

Fungicides. I G FARBERND A-G (Adolf Steindorff, Kaspar Pfaff and Georg Dahmer, inventors) Ger 518,007, Mar 30, 1928 Products comprising finely divided $\text{Cu}(\text{OH})_2$ pptd in the presence of fatty acids of high mol wt. (or their salts) are used Thus, wool fat acids may be dissolved in NaOH , the soln treated with solid CuSO_4 , and the pasty product mixed with tale, dried and ground.

Organo-mercury compounds. T W F CLARK Brit. 338,963, Aug 29, 1929 Phenolphthalein, fluorescein or their homologs or halogen derivs are treated with Hg oxycyanide (or a mixt. of Hg oxide and an alkali cyanide), or reaction is effected with an alkali cyanide on Hg compds of phenolphthalein, fluorescein or their homologs or halogen derivs contg the group HgX in which X represents OH, NO_2 , SO_3H , halogen, etc. The products are bactericides and fungicides Several examples are given

Insecticides. ÉTABLISSEMENTS ET LABORATOIRES GEORGES TRUFFAUT Fr 695,164, Aug 16, 1929 Insecticide solns are prepd by adding as adhesive one or more substances which lower the surface tension of water such as sulfite cellulose lyes, substituted amides of fatty acids or salts of sulfonic acids

Insecticides. UNION CHIMIQUE BELGE. Ger 515,633, Aug 21, 1929 Preps contg the insecticidal constituents of *Pyrethrum cineracholom* are obtained by extg the plant with hydrocarbon solvents, evapg off the solvent and dissolving the residue in pyridine or ethyl lactate Examples describe the extrn of the pulverized flowers with such solvents as pyridine, petroleum ether and alc.

Insecticide and fumigant. CUY H BUCHANAN (to American Cyanamid Co.). U S 1,795,593, March 10. A compn is described which comprises the reaction product formed by fusing together of Ca cyanamid, C and NaF

Pest-destroying composition. PERMUTIT A-G. (Otto Liebknecht, inventor). Ger 515,348, June 15, 1926. Addn. to 472,738 (C. A. 23, 3049) Seed goods, etc., are protected against animal and vegetable pests by a powder contg insol or nearly insol Cr compds instead of the gaseous Cr compds of 472,738 The prepn. may contain $\text{Cr}(\text{OH})_3$, Cr_2O_3 , CrF_3 , CrPO_4 , chromates and bichromates, mixed with talcum, etc.

Composition for protecting plants. CHYVOIN GYÖGYSZER ÉS VEGYÉSZETI TERMÉKEK GYÁRA R. T. (KERESZTY ÉS WOLF) Austrian 120,877, Aug 15, 1930 Plant protecting compns comprising an aq dispersion of S and a hydrocarbon are prepd by stirring a coned aq soln of $(\text{NH}_4)_2\text{S}$ and a hydrocarbon contg an emulsifying agent into a large amt. of water. The hydrocarbon may be petroleum or tar oil, preferably free from phenol, and the emulsifying agent may be Turkey red oil or an NH_4 soap Examples are given.

Preventing infection of plants. ÉTABLISSEMENTS ET LABORATOIRES GEORGES TRUFFAUT. Fr 695,165, Aug 16, 1929 Cryptogamic infections of plants are combatted by the use of org dyes, such as dimethyldiammonaphthopbenasoxonium hydrochloride, mixed with products which lower the surface tension of water

Disinfecting seeds. IGNAZ KREIDL Austrian 121,349, Sept. 15, 1929 Fungicides are applied in paste or pulp form, e g., by mechanically mixing the seeds with the paste or pulp Fungicides sol. or insol in water may be used, and solid diluents, glycerol, etc., may be included in the mixts. Examples are given

Dry-canterizing seed goods. I. G. FARBERND A-G. (Wilhelm Bonrath, Wilhelm Schepss and Carl Taube, inventors) Ger. 515,957, June 7, 1928 The goods are treated with alkyl Hg compds of the general formula RHgX , where R = an aliphatic hydrocarbon residue and X = OH or an acid residue. Examples mention taleum and C mixed with EtHgCl , PrHgBr , BuHgAcO or EtHgAcO

16—THE FERMENTATION INDUSTRIES

C. N. PREY

Formation of kojic acid from sugars by *Aspergillus oryzae*. LADDERICK CHALLENGER, LOUIS KLEIN AND T. KENNEDY WALKER. *J. Chem. Soc.* 1931, 16-23—A strain of *Aspergillus oryzae* was used in these expts in a medium contg inorg salts and designated "medium K." (cf *C A* 23, 4966) Two expts were made with dihydroxyacetone, one contg 11 cc of medium K and 0.5 g of dihydroxyacetone, the other contg 5% dihydroxyacetone. The temp of incubation was 31-2°. Kojic acid was identified in both cases by its m. p. and further characterized in the form of the diacetyl deriv. Kojic acid was also obtained by the fermentation of glycerol by a strain of *Aspergillus oryzae*. Fermentation expts with 5% ethylene glycol, 2% Ca gluconate, 2% KII saccharate, 2% K citrate and 1% glyceric acid were neg. The action of *Aspergillus niger* on glycerol was studied. Trioses were produced when *Aspergillus niger* acted upon a 5% glycerol soln with the salts of medium "M." After 9 days the liquid was filtered, made faintly alk. with ammonia, concd., pptd. with Pb acetate and decompd. with H_2S . Oxalic acid was obtained and identified by titration with 0.02 N NaOH and 0.02 N $KMnO_4$, and by formation of the di-*p*-benzoylphenacyl ester. Neither tartaric nor saccharic acid could be detected. S. JOZSA

Selective fermentation. II. Fermentation of sugar mixtures by Sauterne yeast. HARRY SOBOTKA AND MIRIAM REINER. *Biochem. J.* 24, 1783-6(1930) cf *C A* 25, 166—Sauterne yeast ferments fructose preferably to glucose in a mixt. B. H.

Notes on the determination of cane sugar in sweet wines. H. JAIR. *Z. anal. Chem.* 83, 321-38(1931)—The detn. of sugar in wine was carefully studied and it is pointed out that in detg. cane sugar it is necessary to proceed as in the detn. of the original reducing sugar and use Pb acetate together with $AcOH$. It is recommended, moreover, to simplify the official method of analysis as follows: If the wine contains, x g of ext. per l., take about $3000/x$ g for the analysis. Weigh out the sample accurately in a weighing beaker, transfer to a 100 cc measuring flask and make up to the mark at 15°. Shake and transfer to a somewhat larger dry flask. Add 1.5 g of freshly ignited bone-char and allow to stand 10 min with occasional mixing. For the detn. of the directly reducing sugar, take 25 cc of the filtrate in a 150-cc measuring flask and neutralize with N NaOH. Make up to the mark at 15°, treat 50 cc with Fehling's soln as in the official method. Take 50 cc of the filtrate from the bone-char clarification for the detn. of the cane sugar. Make up to 75 ml in a measuring flask with marks at 75 cc and 100 cc and invert according to the official directions. After the inversion, take 50 cc, dil. to exactly 100 cc, neutralize with NaOH and treat with Fehling's soln as in the detn. of the original reducing sugar. W. T. H.

The acidity of wine and the new (French) regulations. L. FERRÉ. *Ann. fals.* 24, 75-80(1931)—A plea for the adoption of more precise definitions of total, free and volatile acidity, and for the adoption as official of methods capable of yielding results conforming with the definition thus adopted. A. PAPINEAU COUTURE

Detection of fruit wine in grape wine by Werder's sorbitol method. A. SCHREMPF. *Chem.-Ztg.* 55, 52(1931), cf *Chem.-Ztg.* 54, 765, 997(1930)—S states that the cause of contradictory results is due to samples contg. very small quantities of fruit wines. The method of detn. is discussed, the difficulties are emphasized and greater accuracy is suggested. Conclusion: The accuracy and reliability of the method do not depend primarily upon the size of the sample to be analyzed. S. JOZSA

Brandy wines and wine-brandy products. H. ZELLNER. *Chem.-Ztg.* 54, 925-8(1930)—The summary and conclusions of 246 analyses are given. The fusel oil and ester content of brandy wines, distillates and wine brandies are discussed. The figures obtained for yield do not give reliable results in every case, and it is necessary to taste the product for complete estn. The importance of the sense of taste is emphasized and its correlation to various analytical results set forth. S. JOZSA

Infectious turbidity in beer. DE GROENCKEL. *Bull. assoc. élèves inst. sup. fermentations Gand* 32, 117-25(1931)—An address dealing with causes and remedies. A. PAPINEAU COUTURE

Brewery yeast in the chemical industry. FARRER. *Brasserie & malterie* 21, 13-5(1931)—The possibility of production of nucleic acid, amino acids, ergosterol, vitamins and bios from brewery yeast is briefly discussed. A. PAPINEAU COUTURE

The supposed relation between pH value and acid taste of wines (CRISCI) 12. Control and regulation of distilling columns (PÉREZ) 13. The methylene-blue staining

red color obtained on warming it with AcH . The red product is sol in CHCl_3 and 90% EtOH , but insol in Et_2O . A W DOX

Aconitine from Cashmerian *Aconitum chasmanthum*. K H BAUER AND TARA CHAND RADJHAN *Pharm Zentralhalle* 72, 146-52(1931)—An aconitine which differed from all aconitines hitherto reported was isolated from the Cashmerian aconite. The m p and cryst. form are identical with those found for indaconitine, but differs markedly in being insol in Et_2O . W O E

Examination of liquor cresoli saponatus. H NOERDLINGER *Süddeutsche Apoth-Ztg* 71, 159-61(1931)—The method suggested by Kaiser has been subjected to renewed study with the conclusion that it cannot be depended upon to reveal the true compn of cresol soap soln. In the absence, therefore, of a better method, recourse should be had to the official pharmacopoeial procedure. W O E

New drugs and pharmaceutical specialities during the 4th quarter of 1930. F. ZERNIK *Süddeutsche Apoth Ztg* 71, 165-8(1931)—The more important novelties of this period are enumerated and to some extent described. W O E

Percolation or diaecolation. BAEDDIN *Süddeutsche Apoth-Ztg* 71, 172-3(1931), cf C A 24, 4896—The apparent phys and hence economical advantages of the diaecolation method (involving multiple cylinders) over percolation (the so-called American method) in the extn of crude drugs are emphasized and fully discussed. W O E

The reactions of novocaine. L EKKERT *Magyar Gyógyszerésztud Társaság Értesítője* 6, 221(1930), cf C A 24, 4897—A dark red to carmine red color is formed by adding 2 drops of 10% NaNO_2 soln and 10 drops of an alk soln of a naphthol to dil HCl contg a soln of novocaine chloride or nitrate. The soln of novocaine is mixed with 10 drops of CaOCl_2 soln and 0.01 g phenol and NH_4OH added. The mixt becomes green on warming. A soln of chloramine may be used instead of the CaOCl_2 soln. S S DE FINÁLY

Hungarian procedure of morphine manufacture. J KABAY *Magyar Gyógyszerésztud Társaság Értesítője* 6, 226-34(1930), cf C A 24, 3322—Opium poppies are cut up and extd to a jam like pulp of about 0.4-0.8% morphine content. The washing liquid contains free H_2SO_4 which binds Ca and forms easily sol alkaloid salts, prevents the further hydrolysis of chlorophyll and the fermentation or oxidation of the coned ext. The mixt. of alkaloids is then sepd and chlorophyll obtained as a secondary product. S S DE FINÁLY

New reaction for the identification of guaiacol carbonates. K SZÁHLENDER *Magyar Gyógyszerésztud Társaság Értesítője* 6, 235-6(1930)—One-half g guaiacol carbonate is mixed with 0.5 g ZnCl_2 and heated in a dry test tube. A dark brown melt is formed with a very intense smell of guaiacol. The reaction may be brought about faster by using dried or melted ZnCl_2 , since then it is not necessary to wait until the hygroscopic moisture evaps. S S DE FINÁLY

Evaluation of insulin. A STASIAK AND B ZBORAY *Magyar Gyógyszerésztud Társaság Értesítője* 6, 268-74(1930)—Cross tests of Marks (C A 23, 5272) for the evaluation of insulin are modified by injecting into each animal 0.5 cc soln equal to 1 international standard. Three blood tests are taken for sugar detns 15, 3 and 5 hrs after injection. This method gives as exact data as that taking 5 blood tests. The sensitivity reaches 10% in 10-10 expt series, 6 animals being used in each case. S S DE F

Early products in the history of drugs. J HORVÁTH *Magyar Gyógyszerésztud Társaság Értesítője* 6, 292-5(1930)—An address. S S DE FINÁLY

Comparative studies of camomile. P RÖM *Magyar Gyógyszerésztud Társaság Értesítője* 6, 296-9(1930)—The essential oil content of Hungarian camomile was detd (1) by measuring the oil on the surface of the water during distn as carried out in factory practice and adding the oil content of the cohobation water (Results, 0.22 and 0.20%, resp., total oil content, 0.42%), (2) by using the method prescribed by D A B 6 (result, 0.49%). Earlier Hungarian data gave smaller contents since the oil content of the cohobation water was not detd. Some German samples contained 0.6-0.8% a Russian one 0.58%. This latter was a yellow oil rather more sol in water than that of Hungarian or German camomile. Exams of 11 samples of Hungarian camomile of various regions showed oil contents of more than 0.4% for each. The sand content varied from 0.8 to 1.1% and the ash content from 10 to 11%. German samples contained 9-10% ash and 0.4-1.0% sand, Russian samples 11.5-15% ash and 2.7-5% sand. S S DE FINÁLY

Russian flores chamomillae. Á BOROS *Magyar Gyógyszerésztud Társaság Értesítője* 6, 300-6(1930)—Four samples of the Russian drug were examd. It may be distinguished from the Hungarian drug by the presence of *Agropyrum (Triticum) prostratum* L. and *Salsola tamariscina* Paill and further of animal particles of *Amphicoma*

culpes var hirta Falst and *Dulaca Lichatskoi* Humm, which species do not occur in territories east of South Russia. S S DR FINA

The evaluation of commercial Solanaceae drugs. S PERCINI JÓVÁS *Magyar Gyógyszerésztud. Társaság Értesítője* 6, 307-12(1930)—*Atropa belladonna* (I), *Atropa belladonna* (II), *Solanum stramonium* (III) and *Solanum hyoscyami* (IV) of Hungarian origin, examined for alkaloid content according to Ph Helvetica IV. I contained 0.18-0.21, II 0.59-0.66%, III 0.21-0.33% and IV 0.11% alkaloids. Ash and sand contents determined by the method of U S National Formulary, p 428. The values (in %) obtained: ash and sand contents, resp., were I 16.99-17.61, 5.51-5.81, II 6.35-7.52, 0.68-1.11, 18.67-21.80, 2.68-10.00, IV 22.04-22.23, 7.30-7.63. S S DR FINA

Determination of alkaloids, especially the determination of morphine in preparations. E SCHULPK *Magyar Gyógyszerésztud. Társaság Értesítője* 6, 412-51(1930). The indicators used for the titrimetric determination of alkaloids should have pH intervals agreeing with the pH of the alkaloid solution after titration. The morphine determination method Ph Hung 3 cannot be used in the case of some opium preparations (e.g., pantopone drugs). A micromethod was therefore worked out. Mix 0.605 g of opium powder with some water, dilute to 6 cc and filter through dry filter paper. Put 4 cc of the filtrate into a 30-cc bottle, add 0.20 cc NH_4OH and filter the precipitate through dry filter paper. Suck up a part of the filtrate with a micropipette, put into a 30-cc Erlenmeyer flask and add 1 cc purified $PtOAc$ and 0.4 cc NH_4OH . Shake the mixture for 15 min, again add 1 cc $PtOAc$, shake and after 10 min pour the $PtOAc$ solution on a filter paper. Now again add 1 cc $PtOAc$, shake occasionally and pour first the $PtOAc$ solution and the whole liquid on the filter, wash out the flask 3 times with 1 cc water each time. Dissolve the crystalline morphine in 10 cc boiled 0.02 NH_4SO_4 and titrate back the excess acid with 0.02 $N NaOH$. For the evaluation of tinctura opii evaporate 5 g, dissolve residue in 2.5 cc water, add 0.2 cc NH_4OH and dilute to 6 cc. Filter the mixture, 1.5 cc of filtrate into a 30-cc Erlenmeyer flask, add 0.4 cc NH_4OH and determine morphine as above. For the evaluation of extractum opii treat 0.3 g extract with a drop of water and work up as above. The method of Buchbinder (U S Drug Control Laboratory) may well be used for the determination of the morphine content of pulvis Dover. S S DR FINA

Water-soluble oils. K GÖLLNER *Magyar Gyógyszerésztud. Társaság Értesítője* 6, 452-9(1930)—Sulfonation experiments were made on olive, sesame, peanut, rapeseed, nut, and foot oils and on pure white olein. Sulfonated neat's foot oil gives the finest emulsion. It is suitable for cosmetic and skin treating purposes. Cf C A 25, 1000. S S DR FINA

The adsorption capacity of bolus alba (kaolin) for different drugs. Z CSIK *Magyar Gyógyszerésztud. Társaság Értesítője* 6, 460-7(1930)—One to 35 g of bolus alba was treated with 100 cc of different solutions and shaken for 30 min. After sedimentation 10 cc was examined for its content of reagent remaining. 14.4 mg methylene blue was adsorbed from a 1% solution, 0.86 mg $HgCl_2$ from a 1.30% solution, 2.52-11.9 mg I from a 50% alcoholic solution, and 0-0.95 mg nicotine from a 1% solution. No reagent was adsorbed from a 1% solution of phenol or a 3% solution of $HCHO$. Non-starch bolus adsorbed 4.8 water vapor, sterilized bolus 7.2%. The ultraporosity of bolus is small since it has large but not many pores which bind only reagents having low molecular weight. Sterilization often increased adsorptive power. S S DR FINA

Researches on coffeinum natricum, A HIRAY *Magyar Gyógyszerésztud. Társaság Értesítője* 6, 468-73(1930). Experiments showed that 2 mols caffeine are required to form a mol sodium benzoate form an addition compound. The preparation can be made by mixing 11 g caffeine with 20 g ether and 10 g sodium benzoate to pulp and evaporate the ether. The addition of the caffeine and sodium benzoate can be made as follows: Dissolve 0.30 g caffeine in 5 g distilled water, shake with 10 and 5 cc $CHCl_3$, respectively, for 5 min and evaporate the $CHCl_3$ solution at a temperature not higher than 40° . The residue should weigh 0.15 g. Now add 2 cc HCl to the caffeine-free aqueous solution and again shake with 2 portions of $CHCl_3$ for 5 min. Add 30 cc concentrated alcohol to the $CHCl_3$ extract and titrate with 0.1 $N NaOH$ with phenolphthalein as indicator. About 10.3-10.4 cc should be required (1 cc is equivalent to 0.0141 g sodium benzoate). S S DR FINA

Examination of solutions of creosol soaps. I LÉVŐIT *Magyar Gyógyszerésztud. Társaság Értesítője* 6, 474-80(1930)—Five samples of commercial creosol sapon and two samples of home-made liquors (one made according to Ph Hung 3 and other according to D A B 6) were examined for creosol content and for the quality. The quantity of the K soap contained. The fatty acid content and $I-Br$ number were determined. The presence of soap made of oleum jeoriss was proved by the cholesterol reaction of that made of sesame oil, by the furfural reaction, and finally the presence of naphthalene.

colophonic or linolic acids by the CuSO_4 reaction of Charitshkov. The following method was used to det. cresol content. Dissolve 10 g. liquor in a mixt. of 100 cc. water and 30 cc. 10% KOH soln., add 30 cc. 10% BaCl_2 soln., shake the mixt. and dist. to 200 cc. Ba soap ppts. Filter 100 cc. of the purified upper part of soln. contg. alkali cresolates of 0.5 g. of the original liquor. Divide the filtrate into 2 parts, to each of them add 10 cc. dil. HCl and shake with 25 cc. petroleum ether. Evap. the solvent and dry the cresol for 30 min. at 100° . The av. of 2 detns. multiplied by 40 gives the % content of liquor in cresol. 42-50% was found in the samples. The soap examn. may be replaced by the detn. of the emulsifying capacity in the case of cresol soaps of naphtholic and colophonic acids. The cresol content was detd. by the method of Ph. Russ 7 and by steam distn. with CaCl_2 . The former method gave higher values, since colophonic acids are also distd. at the high temp. of the former method and increase the no. of cc. of NaOH used. S S DE FINÁLY

Present state of chemical evaluation of ergot (*Secale cornutum*). P. LIPÁK. Magyar Gyógyszerészeti Társaság Értesítője 7, 9-23 (1931).—The method of Keller-Fromme detns. other materials besides the sp. alkaloids. The method of Forst is exact but too complicated for drug store practice. The D. A. B. method may give inexact data since Mg soaps may be formed and Na_2CO_3 may get into the ether, giving too high a value for alkaloids. Detns. based on the cornutine reaction of Keller-Fromme are inaccurate. A simple acid no. detn. may give some information about the proper storage and age of ergot. S S DE FINÁLY

Determinations of the essential oil content of Hungarian coriandrum. S. PERCSVÁR JÓVÁS. Magyar Gyógyszerészeti Társaság Értesítője 7, 24-9 (1931).—Hungarian coriandrum formerly contained more than 1% essential oil, but in the present harvests it has diminished to 0.6-0.3% content. The sp. gr. diminishes with the increase in oil content. Simple distn. of coriandrum with superheated steam gave higher values for essential oil content than did the method prescribed by D. A. B. This may be explained by the very slight soly. of coriandrum oil in water. S S DE FINÁLY

Colloid-chemical methods in drug stores and laboratories. J. VONDRASEK. Magyar Gyógyszerészeti Társaság Értesítője 7, 59-72 (1931).—An address. S S DE F.

Lycopin. L. CHOLNOVY. Magyar Gyógyszerészeti Társaság Értesítője 7, 95-107 (1931).—Lycopin was extd. from *Tamus communis* and *Solanum dulcamara*. The crude product was purified by recrystn. from CS_2 , petroleum ether. Combustion of this product gave 89.23-89.43% C and 10.63-10.84% H (theoretical compn. of $\text{C}_{40}\text{H}_{56}$, 89.48% C and 10.52% H). The mol. wt. detd. ebullioscopically in CHCl_3 was 520, 564 and 565 (theoretical, 536), that detd. cryoscopically in CHCl_3 , 649, agreeing with Montanari's result obtained in benzene (cf. *Leskione spermatikali agarie uuhane* 37, 909 (1904)). *Tamus lycopin* m. 170° , *Solanum lycopin* m. 174° . No secondary pigments could be found. The data in the botanical literature should be corrected, since the above plants contain lycopin only. S S DE FINÁLY

The volatility of nicotine. L. NAGY. Magyar Gyógyszerészeti Társaság Értesítője 7, 125-30 (1931).—Expts. made with nicotinum puriss. Merck, showed that large losses of nicotine occur during distn. of the solns. Only about 10 cc. of nicotine solns. can be distd. without any appreciable losses in the case of solvents with high b. p. (as petroleum ether, alc.). About 100 cc. can be distd. without significant losses in the case of solvents with low b. p. (ether or a mixt. of ether and petroleum ether). The loss of nicotine by evapn. is greater from beakers than from Erlemeyer flasks. About 0.6 mg. of pure nicotine is volatilized in 5 min. at $20-22^\circ$. S S DE FINÁLY

Senega roots adulterated with sarsaparilla. I. TEMESVÁRY. Magyar Gyógyszerészeti Társaság Értesítője 7, 131-4 (1931).—An ext. of a com. sample of senega roots gave too brown a liquor. No green color was obtained on adding liquor ammonis anisatus but stronger darkening was observed. On examn. under a magnifying glass the sample was found to be adulterated with about 30-5% of sarsaparilla roots. Microscopic examn. showed the structural texture of sarsaparilla roots. The saponin content was detd. by dissolving defibrinized blood with ext. of sarsaparilla particles in physiol. NaCl soln. S S DE FINÁLY

Cascara. T. J. STARKER AND A. R. WILCOX. Am. J. Pharm. 103, 73-97, 147-75 (1931).—A dissertation covering names, species distribution, description, medicinal value, silvies of cascara, growth, collection and curing of the bark, safeguarding the future supply, the industry, est. of existing stand, prospects of future stands and artificial propagation of cascara. A bibliography is appended. W. G. GAESSLER

The analysis of the more commonly used ointments of the British Pharmacopoeia containing an inorganic principle as the active constituent. EOWIN H. BUNCK. Indian

Med. Gaz. 66, 137-8(1931).—Methods are given for the detn. of the base in boric acid ointment, mercury ointment and other ointments that find general employment.

FREDERICK G. GRIMUTH

The potentiometric titration of alkaloids with potassium iodomercurate. LOUIS MAIRIQ *Bull. soc. chim. Belg.* 39, 496-502(1930), cf. *C. A.* 24, 1705, 4895.—The method previously reported for the titration of atropine, hyoscyamine, pilocarpine, euclyonine and sparteine is extended. The accuracy is as good as before. A. L. H.

Chemical composition of certain homeopathic tinctures. JOSEPH S. HERBURN AND ALAN E. SMITH *J. Am. Inst. Homeopathy* 23, 1207-9(1930).—Detn. (as g. per 100 cc.) of (A) total solids, (B) ash, (C) total N, (D) ether ext. of certain homeopathic tinctures yielded the following results: *Dioscorea villosa* A 2.415, B 0.275, C 0.03670, D 0.082. *Hydrastis canadensis* A 1.408, B 0.275, C 0.03123, D 0.161. *Iberis amara* A 2.701, B 0.319, C 0.03902, D 0.164. *Sarracenia flava* (pitcher) A 1.927, B 0.340, C 0.02123, D 0.212. *Sarracenia minor* (pitcher) A 2.205, B 0.094, C 0.03123, D 0.313. (rhizome) A 3.150, B 0.095, D 0.099. *Sarracenia purpurea* (entire plant) A 0.855, B 0.255, C 0.00812, D 0.084. *Sarracenia rubra* (rhizome) A 4.280, B 0.075, *Sarracenia drummandii* (rhizome) A 1.930, B 0.015, C 0.02955, D 0.142. *Tectona grandis* A 0.908, B 0.233, C 0.01078, D 0.100. The presumptive test for the presence of alkaloids was pos. with the tinctures of *Hydrastis canadensis*, *Iberis amara*, *Sarracenia flava* pitcher and *S. rubra* rhizome, and neg. with the other tinctures. The alc. content ranged between 50.15 and 76.62% by vol., and between 49.89 and 69.73% by wt. J. S. H.

A comparison of assays of homeopathic tinctures made according to the American Homeopathic Pharmacopoeia and the Homeopathic Pharmacopoeia of the United States. (A. L. H.). RALPH D. WRIGHT AND JOHN A. BORNLMAN *J. Am. Inst. Homeopathy* 24, 61-70(1931).—Assays for alkaloids were made of tinctures of aconite, belladonna, gelsemium and nuxvomica prep'd according to each of the 2 pharmacopoeias. "The results of the assays indicate that tinctures made according to the Pharm. of the Am. Inst. of Homeopathy are actually 0.1 drug strength tinctures, and that they come up to the requirements of the U. S. Pharm. The results further indicate that the alkaloidal content of tinctures made according to the Am. Homeopathic Pharm. is variable, and that the assay of these tinctures (particularly Class I tinctures) is far below the assumed standard."

JOSEPH S. HERBURN

The determination of cineole in cajaput oil. A. RECLAIRE AND D. B. SPOELSTRA *Ber. Afde. Handelsmuseum Ver. Kolonial Inst. No. 54*, 8 pp (1930).—The method of Cocking (*Perfumery Essent. Oil Record* 18, 165(1927), cf. *C. A.* 21, 3253) was tested on known mixts. of turpentine oil, cineole, terpineol and citronella oil as well as on several samples of cajaput oil from the Dutch East Indies and found to be very useful. Up to 12.5% terpineol does not influence the results, but in case of a higher percentage a correction must be applied which can easily be calcd. from a table given. It is pointed out that the *o*-cresol used in the detn. must be freshly distd., otherwise the results are too low.

J. C. JURJENS

The preparation, composition and value of the syrup of iodotannate. Z. WEINSTECK *Rev. Fac. Cienc. Quím. (Univ. La Plata)* 7, 11-2, 83-95(1930).—There is no loss of I through volatilization in the prep'n. of the syrup according to the formula given in the French Codex Supplement of 1920. Detn. of the I according to the Pharm. of Brazil is simple, rapid and exact as is the method of Ugarte. The entire I in the syrup is in the form of IIt. It is therefore proposed that the syrup be replaced by a simple syrup of IIt, thus eliminating the astringent tannins and their products of oxidation. The prep'n. and assay methods of Machado and Sonol (cf. *C. A.* 24, 5938) are considered unacceptable.

B. S. LEVINE

Determination of moisture in tobacco by means of the hygrometer. N. M. MILOSLAVSKII AND A. T. PALANT *Ukrainskii Khim. Zhur.* 5, Tech. Pt., 117-25(1930).—Expts. show the absence of a simple relationship between the moisture in tobacco and that in the atm. This can be explained on the ground that the amt. of moisture in tobacco is det'd. by its biol. condition and by the adsorptive capacity of the leaf itself. The hygrometer is not applicable to the measurement of moisture in tobacco, because the tobacco leaf is neither chemically nor biologically definite. S. L. MADORSKY

A new method of determining arspenamine preparations. S. KENMATSU AND K. WADA *J. Pharm. Soc. Japan* 51, 65-80(1931).—The sample (0.1-0.2 g.) was digested gently with K_2SO_4 (10 g.), $KClO_4$ (1 g.), water (10 cc.) and dil. H_2SO_4 (40 cc. of 1:1d by vol.). On cooling, 0.5 cc. of 50% soln. of glucose was added and the mixt. digested again. When the soln. became clear it was titrated against 0.05 N $KBrO_3$ soln. with methyl orange as indicator.

F. T. NAKAMURA

Oil of ambrette. R. N. *Parfums de France* 9, 32-8(1931) (in French and Eng-)

hsh) A review of the compn of oil of ambrette and of the constitution of its principal constituents

A. PAPINEAU-COUTURE

The bitter orange and its products. GUIDO ROVESTI *Rivista ital essenze e profumi* 13, 35-61(1931) —From the bitter orange plant are produced the fluid ext., neroli oil, the concrete flower essence, bitter orange water, petitgrain essence, petitgrain water and bitter-orange juice. The fluid ext. is produced by extg the dry powder with 60° alc. It is a limpid green-yellow liquid with bitter taste and pleasant odor, sol in water and sugar syrups. Neroli oil is obtained together with bitter-orange water by distg the flowers with steam. After distn neroli oil is nearly colorless, with time it becomes red with a blue fluorescence. It keeps best in the dark. Ligurian neroli oil has d_{15} 0.8815-0.831, $[\alpha]_D^{20}$ 0.5-6.9, n_D^{20} 1.4710-1.4728, sapon no 28-67.5, esters as linalyl acetate 9.80-23.5%, soly in 80° alc 1-1.2 vols. Calabrian neroli oil has d_{15} 0.727-0.8755, $[\alpha]_D^{20}$ 5.30-6.48, n_D^{20} 1.4715-1.4720, sapon no 52-63, esters as linalyl acetate 17.64-21.56%, soly in 80° alc 1-1.2 vols. Sicilian neroli oil has d_{15} 0.860-1.024, $[\alpha]_D^{20}$ -2.54-5.65, n_D^{20} 1.4680-1.4740, sapon no 6-127. Neroli oil contains pinene 1, camphene I, limonene, dipentene, paraffin, phenylethyl alc., linalool, nerol, terpineol, nerolidol, geraniol, farnesol, decylaldehyde, linalyl acetate, Me anthranilate, benzoates, palmitates, phenylacetates, acetates of different alcs, indole, phenylacetone, a ketone analogous to jasmone, and acetic and palmitic acids. There are 35% of terpenes, 30% of linalool, 18% of linalyl acetate and 1% of Me anthranilate; the ketone gives the perfume. The yield in essential oil depends on the meteoric conditions during the collection of the flowers. The greatest is obtained on warm and dry days. Bitter-orange water is the by product of the neroli oil distn. Its preservation is difficult as it is attacked by the microorganisms. The product has in suspension much neroli oil. The concrete flower essential oil is prepd with petr ether, it is of ointment consistency, having an agreeable odor and coffee color. The yield is 0.230-0.390%. The abs. essence is obtained from the flowers with 4 alc. washings, followed by cooling to -15°, filtration, and evapn of the alc *in vacuo*. The yield is 0.45-0.55%. It has d_{15} 0.0105-0.9228, $[\alpha]_D^{20}$ -0.74 to -5.48, n_D^{20} 1.4658-1.4729, sapon no 105-137.8, esters as linalyl acetate 36.7-48%, soly in 95° alc 1 l by vol. Petitgrain essence is obtained by distg the twigs of the bitter-orange tree. It is yellow, with the odor of the leaves, and contains pinene I, limonene, dipentene, camphene, linalool, geraniol, terpineol, linalyl acetate, geranyl acetate, Me anthranilate and furalol. The Ligurian product has d_{15} 0.890-0.900, $[\alpha]_D^{20}$ -2.5 to -6.7, n_D^{20} 1.460-1.465, sapon no 115-170, esters as linalyl acetate 50-59%, soly in 70° alc 1 vol in 2.5-3 vol. Calabrian petitgrain has d_{15} 0.8971, $[\alpha]_D^{20}$ -0.12, sapon no 214.60, esters as linalyl acetate 75.10%, soly in 80° alc 1 vol in 1.3 vol. The yield is 250-280 g per 100 kg. Petitgrain water is obtained during the distn of the bitter-orange twigs. Bitter-orange juice is obtained from the ripe fruits. It is an intensely yellow liquid, the chem compn of which is not much different from that of sweet oranges, excepting for the bitter principles present. Its principal constituent is limonene (90-97%). Seven hundred g is obtained per 100 kg of fruits. The const. are d_{15} 0.852-0.856, $[\alpha]_D^{20}$ 88-90, n_D^{20} 1.473-1.475, sapon no about 9.

R. SANSONE

Lemon and orange oils. E. BERTÉ *Rivista ital essenze e profumi* 13, 75-7(1931) —A large no of lemon oil samples sponge extd., had the const. $[\alpha]_D^{20}$ 56° 0' to 62° 0', d_{15} 0.8560-0.8590, aldehydes as citral 4.50-6.00%, evapn residue on the water bath 2.50-3.50%, n_D^{20} 1.474-1.4759. Sweet-orange oil had $[\alpha]_D^{20}$ 98° 0' to 100°, d_{15} 0.8485-0.8500, evapn residue 1.50-2.50%, n_D^{20} 1.473-1.475, soly in 90° alc 1.8. In Sicily there exist zones producing lemon oil with d_{15} lower than 0.8560. There is no lemon oil with d_{15} above 0.8500, except that from green fruits, having the low $[\alpha]_D^{20}$ 50° 0' to 54° 0', and a d_{15} sometimes above 0.8500. In 1930 were found lemon oils with $[\alpha]_D^{20}$ 67° 0' and 70° 0' and d_{15} under 0.855. Samples of directly prepd and of distd lemon oil taken in 1931, had n_D^{20} 1.4740-1.467 and 1.4723-1.4745, d_{15} 0.8560-0.8581, citral 3.85-4.45%. In Nov., 1930, orange oil had the av const. $[\alpha]_D^{20}$ 98°, 20' to 100°, 15°, d_{15} 0.8480-0.8498, residue on the water bath 1.50-2.65%.

R. SANSONE

The essential oil of *Satureia nepeta*. L. PEPRINO LIOTTA *Rivista ital essenze e profumi* 13, 93-4(1931) —The oil is a light caramel brown, b 214.5°, showing the absence of terpenic hydrocarbons. The soly in 80° alc confirmed the presence of considerable amts of oxygenated products. The const. were d_{15} 0.9335, $[\alpha]_D^{20}$ 109° 30' soly in 80° alc 1 l, evapn residue 4.82%, sapon no of residue 156. It contained 69% of pulegone. On boiling, the rotatory power rapidly decreases. With 50 min of steam distn all the oil is obtained from the plant.

R. SANSONE

Microsublimation. A. CHALMETA *Anales soc españ fis quim* 28, 1407-19

(1930).—The convenience of microsublimation with caffeine and theobromine and with glucosides is discussed
E M SYMMES

Biphenyl [as source of intermediates for drugs] (MORGAN, WALLS) 25. Alloys [for speculums] (Ger pat 515,798) 9. Emulsions, ointments, etc (Brit pat 339,538) 13.

DUGL F, G Étude sur la dignine. Lons-le-Saunier Deelume 14 pp F 8.
GIBSON, H C Chemistry of Dental Materials. Philadelphia P Blakiston's
Son & Co \$3.50

HADFIELD, CHARLES F Practical Anaesthetics. 2nd ed London Baillière
336 pp 7s 6d, net

Nordiska Specialitetsskommissionens Årsberättelse, 1929 80 pp Analytometoder.
Sanlade och Utarbetade av Nordiska Specialitetsskommissionen Häfte V. Bilag till
Nordiska Specialitetsskommissionens Årsberättelse 74 pp Stockholm Victor Pet-
tersons. Reviewed in J Am Med Assoc 96, 1256(1931)

PLABOT, EMILE Quinquina et quinine Paris Les presses universitaires. 174
pp F 25

WINTERSTEIN, F, AND TRIER, GEORG Die Alkaloide. Eine Monographie der
natürlichen Basen 2nd ed, revised by Georg Trier Teil 2. Berlin Borntraeger

Therapeutic solutions of soporific drugs. I G FARBENING A G Brit 338,983,
Aug 21, 1929 The method of prep'g aq solns of barbituric acids (described in Brit
325,847, C A 24, 4121), by use of a monoalkylated amide of a lower aliphatic acid,
is applied to the prep'n of solns. of other soporific drugs such as the urethan of tri-
chloroethyl alc, tribromoethyl alc, and bromodichloroethyl alc. In examples given,
N-ethylacetamide is used as a solubilizing agent

Medicinal product. E K FREY and HEINRICH KAUT Ger 518,292, Feb 24,
1930 A medicine affecting the circulation of the blood and the action of the heart
is prep'd from urine by freeing this from phosphates in the usual manner (adding uranyl
acetate, filtering off the ppt., decomg the ppt with $(\text{NH}_4)_2\text{HPO}_4$ soln filtering, re-
moving phosphates from the filtrate, and then removing other salts by dialysis. The
final filtrate may be purified by treatment with kaolin or like adsorbent followed by extn
of the adsorbent with $(\text{NH}_4)_2\text{HPO}_4$ soln Alternatively, the adsorbent may be added
to the dephosphatized urine Examples are given

Synthetic drugs. SCHERING-KAHLBAUM A G Brit 339,436, Oct 31, 1929
Dihydroketo-1-alkylhalopyridine compds. are prep'd by treating 2-hydroxypyridine
derivs. contg a halogen atom in the 3- or 5-position (or both) with a simple alkylating
agent or with one which contains a carboxy group (preferably in the presence of an
acid binding substance) Examples are given of the production of 1,2-dihydro-1-
methyl-2-keto-5-iodopyridine, 1,2-dihydro-2-keto-5-bromopyridine-1-acetic acid, 1,2-
dihydro-2-keto-3,5-dibromopyridine-1-acetic acid, 1,2-dihydro-2-keto-5-iodopyridine-1-
acetic acid, 1,2-dihydro-2-keto-3,5-diiodopyridine-1-acetic acid and 2-hydroxy-3-
bromo-5-iodopyridine. The products are sol in water.

Synthetic drugs, etc. I G FARBENING A-G Brit 339,359, Aug 14, 1929
Basic compds. are prep'd by condensation of a sat'd or unsat'd fatty acid contg at least
10 C atoms, or an ester, chloride, amide or imide ester of such an acid (such as palmitic
or oleic acid chlorides) with an aromatic primary or secondary amine having a side
chain contg tertiary N such as β - β -diethylaminoethoxyaniline, β -aminobenzodimethyl-
aminoethanol and unsym ethyl[or methyl]diethylamine ethyl- β -phenylethylamine

Synthetic drugs. I G FARBENING A G (Karl Streitwolf, Alfred Fehle, Paul
Fritzsche and Walter Herrmann, inventors) Ger 518,208, June 16, 1929 Acylamino-
benzenestibinic acids substituted in an α -position to the stibinic acid residue are prep'd
by the customary processes. Thus, m - or p -acylaminoanilines substituted in the α -posi-
tion to the NH_2 group, e g. with OH , OCH_3 , Cl , or halogen, and free from or contg
substituents, may be diazotized and then treated with an antimonite Alternatively,
 m - or p -aminobenzenestibinic acids substituted in the α -position to the acid residue, and
free from or contg other substituents, may be acylated Examples are given The
products are effective against diseases caused by trypanosomes, and may be adminis-
tered per os

Alkaloids. GEORGE KNOTT Ger 516,283, Nov 28, 1928 Addn to 497,630 In
obtaining volatile alkaloids by steam distn by the methods of 497,630 (C A 24, 4121),
500,294 (C A 25, 560) and 500,998 (C A 24, 4900), the distn takes place at a steam
pressure of above 2 atm The process is especially applicable to nicotine, a mash of

tobacco being treated with a base such as NaOH to free the alkaloid before distn. In the example the nicotine yield is increased from 0.16% to 2.12% by increasing the steam distn. pressure from a pressure of 0.5 atm. to 5 atm.

Caffeine. KAFFEE HANDEL A. G. Ger 515,604, Oct 23, 1923. Caffeine is extd from raw coffee beans by moistening the beans and extg with CCl_4 at low temps.

Aqueous solutions of pure glucosides. G. RICHTER, LTD. Hung 101,049, June 28, 1930. Glucosides sepd from various plants or drugs are made water sol by adding salts or mixts of cholic acid or its deriva.

Anthraquinone glucosides from drugs such as cascara and frangula. FRIEDRICH KAYSER and KARL SCHRANZ (to Winthrop Chemical Co.) U. S. 1,706,109, March 10, 1928. The water sol anthraquinone glucosides may be isolated in substantially pure form by a procedure involving the treatment of alc. or other exts of the drugs with $\text{Fe}(\text{OH})_3$, Mn hydroxide or Al hydroxide, thereby sepg substantially all of the inert material, and subsequent drying of the solns of the glucosides so isolated. The dried products, thus obtained in the form of fine leaflets, represent the substantially pure effective anthraquinone glucosides or mixt of glucosides contained in the drugs. They are sol in water with a slightly acid reaction forming clear solns generally of a brownish color, adaptable for use as purgative injections. Examples with details of treatment of cascara and frangula are given.

Betaine thiocyanate. I. G. FARRBEND A. G. Ger 515,545, June 29, 1927. See Brit 316,693 (C. A. 24, 1937).

1-(*m*-Aminophenyl)-2-methylamino-1-propanol. FLEMA E. MEACK (Otto Dalmer and Max Oberlin, inventors) Ger 518,212, May 16, 1929. The customary reduction methods are applied to 1-(*m*-nitrophenyl)-2-methylamino-1-propanone. Examples are given. The therapeutic properties of the product compare favorably with those of 1-phenyl-2-methylamino-1-propanol. Cf. C. A. 25, 1264.

Basic tertiary alcohols. I. G. FARRBEND A. G. (Fritz Mietsch, Josef Klarer and Hans Hahl, inventors) Ger 518,211, Mar 3, 1928. Compds in which an amino or alkylamino group is linked by an aliphatic residue to an aromatic, hydroaromatic, or heterocyclic ring attached to a tertiary alc. group, are prepd by the action of org. Mg compds on esters of cyclic carboxylic acids contg an aminoalkyl or alkylamino-alkyl residue directly substituted in the nucleus or attached thereto through N or O. Thus diphenyl *p*-(diethylaminoethylamino) phenyl carbinol may be prepd by treating ethyl *p*-diethylaminoethylaminobenzoate with the Grignard reagent from Mg and PhBr. Other examples are given also. The products are therapeutically useful and resemble ephedrine, hardenine and adrenaline in their effects.

Hydroxyanthraquinone derivatives. WALTER STRAUB Ger 518,215, Feb 6, 1929. The total content of hydroxyanthraquinone deriva, both free and combined, is extd from drugs contg them, e. g., cortex frangulae, by means of a moist org. solvent not completely miscible with water, e. g., Et_2O , C_6H_6 or $\text{C}_6\text{H}_5\text{Cl}$. An example is given.

Gold compounds of succinimide. W. J. FORD Brit 338,500, July 19, 1929. Therapeutic compds (various formulas of which are given) are prepd by the interaction of chloroauric acid or Au hydroxide or a salt thereof, or fulminating gold, with succinimide in the presence of a base (suitably by heating in alc.). Compds contg Br and the CNS radical also are mentioned.

Increasing the hypoglycemic action of insulin. G. RICHTER, LTD. Hung 101,262, Jan 28, 1930. Insulin is mixed with ext. of tonilla.

Antrachinone preparation from yeast. GY. FENES Hung 100,690, Aug 4, 1927. Yeast is boiled with water and KOH or NaOH, then mixed with org. solvents. The upper phase is sepd, coned and cooled. The fluffy crystals formed are recrystd from alc., dried and irradiated in thin layers by a mercury lamp. The product is incorporated in cacao butter, mixed with sugar and pressed into tablets.

Vitamin-containing products. AAGE W. OWSE Ger 472,814, Oct 11, 1924. See Norw 44,018 (C. A. 25, 1336).

Liver esterase. HEINRICH KRAUT Ger 518,024, Jan 5, 1928. Liver exts are dialyzed against running water, whereby salts are removed and part of the albumin content is flocculated. The residual soln is treated with an adsorbent, e. g., kaolin or Al_2O_3 , which adsorbs the remaining albumins. The purified solns so obtained are suitable for therapeutic use. An example is given.

Hormones. SIGMUND FRANKEL Austrian 121,007, Aug 15, 1930. Details are given of the extn of hormones from testes by a process similar to that described in Brit 292,962 (C. A. 23, 1474-5). Cf. C. A. 24, 3608.

Active lipoids. HENRI ISCOVESCO (to Health Products Corp.) U. S. 1,796,027, March 10, 1929. Dry pulverized organs such as testes, heart muscle, brain, adrenal cortex,

adrenal total, intestine, stomach, ovary, corpus luteum, red corpuscle, mammary gland, thymus, kidney, pancreas, total pituitary, placenta, prostate, lung, spleen, thyroid and liver are extd with ether, CHCl_3 and alc. The solns thus obtained are treated with acetone to effect pptn of a lipid contg material, and the ppt is sepd from the soln. and is treated successively with cold abs. alc., ether and CHCl_3 to effect purification if desired by soln and subsequent reprecipitation with acetone. The starting material, such as cod livers, may be given a preliminary extn with acetone. Various details and modifications of procedure are described.

Purifying olive oil or other vegetable oils. H. HATAKEYAMA and H. WATANABE. Brit. 339,011, Sept. 12, 1929. Remaining traces of fatty acids in partially purified oil are esterified with an aromatic alc. such as benzyl alc., phenylethyl alc., hydroxybenzyl alc. or cinnamic alc. to prepare a fully purified oil suitable for use as a solvent for medicinal injections.

Sterilization vessel for injection liquids. RUDOLF SCHMITZ. Ger. 515,947, Mar. 19, 1929.

Soaps, creams, etc. AUBERT C. J. PARBENT. Fr. 695,218, Aug. 28, 1929. Esters of the *p*-aminobenzoic and phthalic series are incorporated in soaps, creams and other toilet preps. Those esters classed as toxic are excluded.

Vermicidal composition. WM. E. WATERHOUSE. U. S. 1,796,070, March 10, 1929. A compn. suitable for killing worms consists of a water-sol. brown powd. material prepd. from powd. oil-extd. mowrah seeds by first extg. with water and then evapg. this soln. to dryness.

Improvement of tobacco and tobacco products. I. SCHÖN and I. HEACZECH. Hung. 100,779, Sept. 24, 1929. Tobacco is moistened with a soln. contg. hygroscopic disinfectant and must preventing agents, e. g., 1000 g. water, 2.5 g. camphor, 1 g. saffron flour, 100 g. sirupus kaliumsulfogajacoeus, 100 g. dild. alc. and 6 oz. ZnSO_4 .

18—ACIDS, ALKALIES, SALTS AND SUNDRIES

E. M. SYMMES

Manufacture of synthetic nitric acid and its salts. E. GUERIN. *Riv. chim. ind.* 39, 258-62 (1930).—A review. P. THOMASSET.

The process of intensive production of sulfuric acid. E. OWSIANY. *Chem.-Ztg.* 55, 76 (1931).—App. is inserted between Glover towers and the first chamber, and between the first 2 chambers, mechanically to disperse, compress and expand the components, to ensure better contact. The consumption of pyrites was increased from 36.3 to 46.8 tons, and the H_2SO_4 made in the chambers decreased 11.8%, the balance being produced in the intensive mixers. E. M. SYMMES.

The navy's sulfuric acid plant. H. M. COSTER. *U. S. Naval Inst. Proc.* 57, 317-9 (1931).—This plant, at the Naval Powder Factory, Indian Head, Md., is of the Grillo-Schroder type with Pt-MgSO_4 contact mass, having a rated capacity of 70 tons of 100% H_2SO_4 per 24 hrs. Several special features of the plant are described: (1) a SO_2 recorder operating on thermal-cond. principles, which enables the operator to supply gas of uniform SO_2 content to the converter, (2) a special app. for detg. the strength of oleum by cond. measurement; (3) a unique method for adding automatically the necessary water to the absorbing towers by admitting a mixt. of water and steam into a by-pass line in the circulation system. These features greatly reduce the cost of operation and simplify plant control. C. G. STORM.

Problems of a Hungarian nitrate factory. T. SZAFKA. *Technika* 10, 155-60, 185-91, 247-52 (1929).—Methods of nitrate manuf. in different countries are described. The Hungarian factory should avoid arc fixation of N_2 , and should only use Ca cyanamide or synthetic NH_3 methods. S. S. DE FRIALLY.

The manufacture of potassium metabisulfite. I. FERRAN. *Quim. e ind.* 7, 277-9 (1930).—The manufacture of $\text{K}_2\text{S}_2\text{O}_5$ is described. Cost figures (in pesetas), in a plant making $\text{Na}_2\text{S}_2\text{O}_5$ at the same time, are 1027 kg. of 92% KOH 1129.7, 1422 kg. of 98% K_2CO_3 1707.6, 1138 kg. of 98% S 318.65, 247 kw. hrs. 74.1, water and light 10, 1400 kg. of Cardiff coal 112.0, salaries 135.3, 27 workmen 189, watchman 50, amortization 44.45, charges 71.7; metabisulfite obtained 3100 kg.; cost per 100 kg. of metabisulfite 117.4. S. L. B. ETHEXTON.

Outline of a potassium dichromate plant. VICTOR MALDA. *Chimica e industria* 3, 470-4, 510-3 (1928).— $\text{K}_2\text{Cr}_2\text{O}_7$ is obtained by an oxidation of chromate or Cr ore in the presence of CaO and Na_2CO_3 . The app. required and a plant sketch, with a calcn.

of cost for a daily production of 3000 kg $K_2Cr_2O_7$, is given in detail. In Brazil $K_2Cr_2O_7$ can be produced for 876 reis per kg. R. D. BUMBACHER

The production of potassium sulfate from potassium chloride and sulfur dioxide. E. BRITZKE, S. VOLPKOVICH AND M. KAMENSKAYA. *Udobrenie i Urozhai (Fertilizers and Crops)* 2, 492-5 (1930).—Expts with a no. of catalytic agents to speed up the reaction between KCl and SO_2 in the presence of water vapor were made. At 500°, without a catalyst, only 4.5% of the KCl decompd. in 20 hrs. $CoSO_4$, $Co(NO_3)_2$, Fe_2O_3 , $Fe(OH)_3$, Al_2O_3 , CuO , burned pyrites and a series of slags contg oxides of Al, Fe, Si, Ca and Na were tried. These were prepd. in the lab. or taken from com. plants. Max. effects were obtained from the slags contg oxides of Fe, Al, Si and Na. The slag was ground to pass a sieve with 4000 holes per sq. cm., and 0.5% of the slag was added to the salt, which was ground to particles of 0.2-0.4 mm. diam. The yield of sulfate reached 98.2-99.8% in one hr. With 0.1% Fe_2O_3 or 0.01% $Fe(OH)_3$, (of the wt. of KCl) the yield was 78-83% in 3.75 hrs. Burned pyrites gave a yield of 70% in 2 hrs. when 1% was used. With 0.5% CuO the yield reached 44% in 20 min. Various other slags were tried out, but without much improvement. J. S. JOFFE

Formation of carbon disulfide from hydrogen sulfide and coke. T. J. DRAKELEY AND W. J. BAKER. *J. Soc. Chem. Ind.* 49, 475-8T (1930).— CS_2 in coal gas may be formed by decompn. of org. S compds. in the coal, or by reactions between C and free S derived from pyrites. When H_2S was passed over heated wood C or coke, an extended delay occurred before evolution of gas and CS_2 formation, caused by H_2S absorption by C. Small quantities of CS_2 were formed by the action of H_2S on wood C between 700° and 1000°, or on coke between 600° and 950°. In the tests the max. CS_2 yield was obtained at 2.4 l. of H_2S per hr. passed over. CS_2 was not obtained by the action of H_2S on coke from coal carbonized above 950°, nor was CS_2 obtained from H_2S and C at less than 500° or above 975°. Over 90% of CS_2 is decompd. on passing over wood C or coke at 850-1000°. E. M. SYMMES

Titanium oxide. RENÉ SPITTEL. *Riv. prod. chim.* 34, 1-5, 40-4 (1931).—A review. P. THOMASSET

The kinds of lime and their manipulation. F. EISEMANN. *Tonind.-Ztg.* 55, 233-40 (1931). F. O. ANDEREGG

Quinquennial review of the mineral production of India for the years 1924 to 1928. Saltpeter. E. H. PASCOE. *Records Geol. Survey India* 64, 287-92 (1930).—The Gangetic plain is famous for its production of saltpeter which is extrd. from the earth (1-25% salt). Wood ashes are added to decompose $Ca(NO_3)_2$. The liquors analyze NaCl 15.25, KNO_3 7.24, KCl 0.40, $MgCl_2$ 0.20, $CaCl_2$ 0.10, $CaSO_4$ 0.10 and $MgSO_4$ 0.10%. ALDEN H. EMERY

The solubility of salt, alone or in combination, in water and brines at temperatures below and above 100°. W. FROEHLICH. *Mitt. Kalk-Forschungs-Anst.* 1929, 37-40, *Chem. Zentr.* 1929, II, 2485.—F. detd. the solubilities of KCl, NaCl and $MgSO_4$ in H_2O , of NaCl in the presence of KCl in H_2O , of $MgSO_4$ in H_2O contg. different quantities of $MgCl_2$, of NaCl in the presence of KCl in technical sylvinite extrn. liquor and in van't Hoff's liquor Q 20, and of carnallite in the liquor Q 20. All detns. were made in an autoclave between 100° and 200°, in a few cases also some known figures at temps. below 100° were reexamd. The values obtained are plotted in wt. %, and the curves are continued to the fusion temp., where the relation of salt to solvent becomes 100/0. For quantities of KCl + NaCl dissolved simultaneously in 100 g. soln. this line is absolutely straight, and at 722° it meets the m. p. of a mixt. contg. 1 mol. NaCl + 4 mols. KCl, according to the known fusion diagram of the mixt. of both chlorides, while from 120° downward NaCl shows an unvaried soly of 16.4%. The most recent detns. of the soly of $MgSO_4$ were confirmed, and deviating results obtained by F. or earlier investigators were traced to the presence of sediments consisting of labile hydrates, these hydrates hold an intermediate position between the stable hydrates $MgSO_4 \cdot 6H_2O$ and $MgSO_4 \cdot H_2O$. G. SCHWOCK

The processing of sylvinite at temperatures above 100°. Dissolving, crystallizing and heat consumption. W. FROEHLICH AND E. RYTER. *Mitt. Kalk-Forschungs-Anst.* 1929, 67-69, *Chem. Zentr.* 1929, II, 2485.—By means of the figures obtained for soly in solns. satd. with NaCl and KCl (cf. preceding abstract) the sepn. of dissolved salt on cooling the soln. is calcd. The soly. of NaCl in a soln. satd. with KCl is greatest at 80°. Thus when such a soln. is cooled, NaCl + KCl is sepd. down to this temp.; NaCl is just dissolved again by solns. satd. at 150° and cooled to 20°, provided that the solns. are thoroughly agitated. Thus at a soln. temp. below 150°, KCl free from NaCl can be produced, while above 150° only KCl contg. NaCl is obtained. When the cooling is accompanied by evapn. of H_2O , as occurs especially in the vacuum cooling app.,

a mixt. of NaCl and KCl is obtained, which affects the results accordingly. In any case, a quantity of H_2O sufficient to keep the NaCl in soln can be added. The kind and proportions of salts and liquor, as well as the H_2O to be added, are calcd. for a dissolving temp. of 100° and 200° , and for cooling with or without evapn and with or without addn of H_2O . These calcns. are carried out with reference to the practical conditions met with in the soln. of sylvinit. These conditions are discussed, together with thermotheroretical considerations on soln. at temps. above 100° . G S

Modern arrangements for extracting salt from the ocean in East Africa. Fa. RIEDIG. *Intern. Bergwirtsch. u. Bergtech.* 23, 142-3 (1930), *Kali* 24, 271-3, 294-7 (1930). — A brief description of the salt-extrn. plant in the Italian colony Somaliland, claimed to be the largest ocean salt works in the world. E I S

Commercial manufacture of hydrogen. E GUEVIN. *Rev. chim. ind.* 39, 354-7 (1930). — A short description is given of the methods in use. P THOMASSET

Mechanical mining and treatment of sulfur. J B NEALEY. *Eng. Mining J.* 131, 253-5 (1931). — The practices at the Hoskins Mound and Brynmound, Tex., plants of the Freeport Sulphur Co. are described. W H BOYNTON

Mica. J H FAYBLENDER. *Rev. prod. chim.* 33, 549-54, 584-6, 616-9, 654-5 (1930). — A review. P THOMASSET

Bleaching earths. O ECKART. *Seifen- u. Ztg.* 58, 132-3 (1931). — Kaolins with out bleaching properties yield their H_2O at $450-700^\circ$, while bleaching earths give off their H_2O rather uniformly between 200° and 900° , the nearer this curve comes to a straight line, the better is the bleaching power. A bleaching earth that has been acid treated follows the same rule, this indicating that bleaching earths possess a gel body and non-bleaching earths a cryst. body. By mixing 10 g. of oil with 10 g. of earth at room temp. bleaching earths showed an increase in temp. (heat of wetting) of $4-13^\circ$, while non bleaching earths showed a rise of 1° . P ESCHER

Washing and bleaching clays of Azerbaijan (bentonites and floridines). S A KOVALEVSKI. *Azerbaidzhanskoe Neftyanoe Khozyaistvo* 50 pp (1931). — Natural deposits are described. Bleaching clays as yet discovered require activation with H_2SO_4 , but scarcity of acid prevents their immediate utilization. V K

Preparation of HCN (GLUUD, KELLER) 6. Apparatus for mixing acids (Brit. pat. 338,410) 1. HNO_3 production (Brit. pat. 338,497) 13

Kab-Kalender, 1931. Taschenbuch für Kalibergbau und Kalindustrie. Edited by C. HERMANN. Halle (Saale). W. Knapp. M 5.20

KAUSCH, OSKAR. Die Kontaktstoffe der katalytischen Herstellung von Schwefelsäure, Ammoniak und Salpetersäure. Halle (Saale). W. Knapp. 216 pp. M 21, linen, M. 23

LENDEL, E. Interferometrische und spektroskopische Untersuchungen zum Nachweis von Unterschieden zwischen natürlichen Quellsalzen und ihren künstlichen Ersatzprodukten. Berlin. Schoetz. 39 pp. M. 3

STANDAGE, H C. Cements, pastes, glues and gums. London. Crosby Lockwood. 164 pp. 3s 6d

Chromic acid and sodium bisulfate. HAWSHAW CHEMICAL CO. Brit. 338,938, July 30, 1929. Reaction between oleum and $Na_2Cr_2O_7$ is effected in an app. (details of which are described) so that molten products are produced directly which are then sepd. by centrifuging or settling; either or both products may be solidified by spraying or atomizing.

Nitric acid. N CARO and A R FRANK. Brit. 338,566, Aug. 13, 1929. NH_3 is oxidized with O or gases contg. O under normal pressure, resulting N oxides are compressed to "several atm." pressure in a turbo compressor and at this pressure are converted into HNO_3 by condensation and absorption. Cf. C. A. 24, 594b

Sulfuric acid. WILLI BÜSCHING. U S 1,795,995, March 10. See Fr 673,105 (C. A. 24, 2552).

Concentrating sulfuric acid. P I KHAZANTOV. Russ. appl. 58,870, Nov. 23, 1929. H_2SO_4 is coned. by hot vapors from the Kessler and Gaillard app. by first directing these gases contg. SO_2 and SO_3 into a superheater, where they are heated to $900-1200^\circ$, followed by their recycling through the concn. app. to concn. new portions of H_2SO_4 , and to eliminate losses in S oxides by means of a continuous recycling of gases used for heating.

Material resistant to acids and alkalies. KÄNELGÅR, LTD. Hung. 101,471, Jan.

24, 1930. S and Se are used together with animal hairs, paper sheets, waste cloth or materials which in themselves are not resistant to acids and alkalis. The mixt. is pressed to forms then heated under pressure to a temp. higher than the m. p. of S or of Se.

Alkali and alkaline earth metals. JAN H. DE BOER (to N.-V. Philips' Gloeilampen fabriek) U. S. 1,797,131, March 17. See Brit. 323,718 (C. A. 24, 3229).

Alkali and alkaline earth carbonates and chlorides. SOC. ANON. ALCALINA. Brit. 339,330, Aug. 1, 1929. See Belg. 358,292 (C. A. 24, 925).

Alkali chromates. BOZEL-MALETRA (SOC. INDUSTRIELLE DE PRODUITS CHIM.) Fr. 37,257, Jan. 28, 1929. Addn. to 683,604. The amt. of alkali used is limited to the equiv. necessary for the formation of dichromate instead of chromate.

Alkali chromates. O. F. TARR (to Mutual Chemical Co. of America) Brit. 338,469, April 3, 1929. See U. S. 1,752,863 (C. A. 24, 2553).

Alkali chromites. BOZEL-MALETRA (SOC. INDUSTRIELLE DE PRODUITS CHIM.) Fr. 37,256, Jan. 25, 1929. Addn. to 683,179 (C. A. 24, 4593). Alkali chromites are prepd. by heating ores of Cr in the presence of alkalis in a reducing atm. or in a reducing medium, with or without pressure. The Fe obtained is sepd. by known means.

Alkali and alkaline earth metal cyanates. DEUTSCHE GOLD- UND SILBER-SCHNEID-ANSTALT VOM ROESSLER. Brit. 339,220, April 10, 1929. Urea is heated in the absence of solvents (suitably to 120-150°) with basic alkali metal compds. such as soda or potash or alkali hydrides (but excluding alkali hydroxides) or with alk. earth metal compds. such as CaO, Ca(OH)₂, or Ca carbide, hydride, amide or carbonate. NH₃ is evolved and recovered, and C₂H₂ also is formed when CaC₂ is used. Alk. earth metal cyanates obtained may be treated in aq. suspension or soln. with alkali metal compds. such as NaCl to produce alkali metal cyanates.

Cyanates and cyanamides. I. G. FARBENIND. A.-G. Brit. 339,371, Sept. 13, 1929. Cyanates and cyanamides are prepd. by heating urea with oxides or carbonates of bivalent metals such as the alk. earth metals (including Mg) and metals above Hg in the electromotive series such as Zn, Cd, Cu, Ni and Pb. At temps. between 130° and 400° mainly cyanates are formed, and at higher temps., cyanamides. NH₃ and CO₂ can be used instead of urea, and various modifications and details of procedure are described.

Cyanides and ferrocyanides. EMIL HEINE. Ger. 315,850, Dec. 22, 1928. Alkali sulfates, CaO and H₂S are allowed to react to produce alkali hydrosulfides and CaSO₄. The latter is removed by filtration. The hydrosulfides are obtained by evapn. and heated with C₂N₂(NH₄)₂ to produce alkali thiocyanate and NH₃. The thiocyanate is heated with Fe or H or both to produce alkali cyanide and FeS or H₂S. The FeS reacts with the alkali cyanide to form ferrocyanide. Extn. with cold water removes the alkali cyanide, and extn. with hot water removes the ferrocyanide and alkali sulfide, which are sepd. by fractional crystn. The alkali sulfide is worked up to alkali sulfate.

Alkali metal hydrides. HERMANN FREUDENBERG and HARRY KLOEFFER (to Deutsche Gold und Silber-Schneideanstalt vom Roessler) U. S. 1,796,265, March 10. Finely divided alkali metal is subjected to the action of H at temps. of about 180-300°.

Alkali phosphates; hydrogen chloride. METALLGES. A.-G. Ger. 518,203, Dec. 30, 1928. These are prepd. from alkali chlorides and P₂O₅ by introducing vaporized or solid P₂O₅ into a melt of alkali chloride into which steam is led.

Alkali sulfates. KALI-CHEMIE A.-G. Ger. 515,930, Jan. 18, 1927. In producing alkali sulfates from alkali chlorides by treating them with flue gases and steam, a 1% addn. of surface active material such as clay, bauxite, SiO₂, etc., is present. Thus, the chlorides may be briquetted with rock salt, burned clay or fuller's earth before submision to the flue gas and steam.

Aluminates. MAX PASCHKE. Ger. 518,204, Feb. 15, 1927. Water-sol aluminates are obtained as by products in the manuf. of crude Fe, by adding Na compds., particularly Na₂CO₃, to the blast furnace charge. The aluminates are extd. from the slag with hot water.

Metal salts by the double decomposition of salts with sulfates. I. G. FARBENIND. A.-G. (Otto Balz, inventor) Ger. 514,651, Nov. 15, 1927. HPO₄ or its salts are added. Gypsum is obtained as a by product if Ca salts are used. Thus, Ca(NO₃)₂ soln. is treated with HPO₄ and hot 90% K₂SO₄ soln. The resulting KNO₃ soln. is readily sepd. from the coarse CaSO₄ produced. Other examples are given.

Metal carbonyls. I. G. FARBENIND. A.-G. Fr. 37,284, July 29, 1929. Addn. to 677,548 (C. A. 24, 3329).—The process of the prior patent is carried out by mixing the primary material with liquids or molten masses and afterward treating with CO.

to produce carbonyls under such conditions that the liquids or molten masses are in the gaseous phase or almost exclusively so during the reaction.

Aluminum chloride. I. G. FARBENIND. A-G. Fr. 693,124, May 6, 1930. See Ger. 515,033 (C. A. 25, 1644).

Ammonium nitrate. E. E. LIDER and I. E. KOROSCHANSKI. Russ. appl. 35,845, Nov. 22, 1928. NH_4NO_3 is prepd. from HNO_3 and NH_3 in a vacuum neutralizer connected by pipes successively with a column which acts as a trap for NH_3 and with a heat exchanger used for preheating HNO_3 passing into the neutralizer.

Ammonium nitrate. ODDA SMELTEVERK AKTIESELSKAPET and E. JOHNSON. Brit. 339,562, March 4, 1929. See Fr. 690,757 (C. A. 25, 1542).

Ammonium sulfate. I. BALINT. Hung. 101,900, May 26, 1930. Na_2SO_4 is treated in soln. with NH_3 and CO_2 . NH_3 , however, is not regenerated from the mother liquor after the pptn. of NaHCO_3 , but evapn. is continued and the soln. cooled until $(\text{NH}_4)_2\text{SO}_4$ is produced. To mother liquor contg. $(\text{NH}_4)_2\text{SO}_4$ may be added NH_4NO_3 in order to produce the double salt for use as fertilizer.

Borax. AMERICAN POTASH & CHEMICAL CORP. Fr. 37,316, Aug. 10, 1929. See Brit. 330,453 (C. A. 24, 5949).

Carbon disulfide. ZAHN & CO. G. M. & H. Fr. 37,206, July 10, 1929. Addn. to 672,766 (C. A. 24, 2352). The CS_2 is cooled in the condenser to a temp. a little below its b. p. so that only a small quantity of H_2S is absorbed, which is afterward removed in known manner in the separator. Preferably, the gases liberated from the condenser are brought into a second condenser of the same kind but of smaller dimensions, from which the condensed materials are returned to the retort.

Separating calcium and magnesium chlorides from brines. WM. R. COLLINGS (to Dow Chemical Co.). U. S. 1,790,920, March 17. Tachydrite crystals are sep'd. from a mother liquor having a higher proportion of CaCl_2 to MgCl_2 than the original soln.; the crystals are dissolved in a limited quantity of water with heating such that a soln. is produced from which on cooling to about 30° crystals of MgCl_2 hexahydrate will be pptd. and the filtrate remaining from the crystals will be of approx. 42.5 – 43° B ϕ . This filtrate retains the CaCl_2 in soln. at 30° .

Calcium-magnesium chloride. WM. R. COLLINGS and JOHN J. SILVER (to Dow Chemical Co.). U. S. 1,790,883, March 17. A free-flowing non-caking mast. of hydrated crystals of chlorides of Mg and Ca is pptd. as crystals of tachydrite and CaCl_2 dihydrate from a soln. satd. with the chlorides, and the crystals are superficially dehydrated to the point which avoids caking together. U. S. 1,791,886 relates to superficially dehydrating crystals after pptn. as crystals of the hydrated double chloride $\text{MgCl}_2 \cdot 2\text{CaCl}_2 \cdot 6\text{H}_2\text{O}$ and CaCl_2 dihydrate. Cf. C. A. 24, 927.

Calcinng lime or dolomite. W. VOSS and VULCAN-FABRIKUNG A-G. Brit. 339,103, Nov. 27, 1929. Water which is injected into a shaft kiln for calcining lime or dolomite is cooled nearly to the f. p. to prevent its conversion into steam before it contacts with the charge.

Hydrating lime. CLARENCE R. REX. U. S. 1,796,396, March 17. Various details are described for mixing a measured quantity of water with a measured quantity of quicklime, detg. the vol. of the water and the rate of supplying it to the lime and mixing to approx. a constant temp. during the reaction (the reaction being checked as complete by predtd. increase in vol.) so as to constitute a standardized com. operation.

Copper sulfate briquets. HUNGÁRIA MŰTÁRGYVÁR, LTD. Hung. 101,213, Feb. 1, 1928. CuSO_4 crystals of 0.25 to 5.00 mm. are formed by interrupted crystn. and briquets are formed from these. Generally no cementing matter is required since postcrystn. takes place after briquet formation. The briquets may be treated with concd. mother liquor.

Hydrogen peroxide. ELEKTROCHEMISCHES WERKE-MÜNCHEN A-G. Ger. 515,596, Sept. 7, 1926. Solns. of H_2O_2 in caustic alkali are preserved and used in vessels of Al or its alloys.

Ferric phosphate. METALLGES. A-G (Freiherr Conway von Gerswald and Hans Weidmann, inventors). Ger. 518,315, May 30, 1929. The manuf. of FePO_4 by roasting ferro-phosphorus in air is accelerated by addn. of a catalyst, e. g., NaCl or Na_2CO_3 .

Iron carbonyl. I. G. FARBENIND. A-G. (Alwin Mittasch and Carl Müller, inventors). Ger. 518,387, Dec. 16, 1924. Addn. to 499,291 (C. A. 24, 4127). The method of Ger. 499,296 is extended to the use of Fe in non-spongy form. The preliminary reduction, if applied, should take place in a vessel sep. from that in which the carbonyl is prep'd.

Liquefying nitrogen oxides. I. G. FARBENIND. A-G. (Christoph Beck and Heinrich Diekmann, inventors). Ger. 518,386, Aug. 19, 1927. In sepg. N_2O_4 from gas

mixts with or without other oxides of N, by liquefaction at atm. or raised pressure, the vapor pressure of the N_2O_4 is reduced by addn. of up to 10% of HNO_3 in liquid or vapor state. The added acid may first be enriched in oxides of N. If the gas mixt. is moist and contains O, the necessary acid may be formed *in situ*. Cf. C. A. 24, 670.

Apparatus for adsorbing nitrogen oxides. V. I. MALYAREVSKII and V. V. PAPKOV. Russ. appl. 21,940, Nov. 28, 1927, 37,713 and 37,714, Dec. 28, 1928. Mechanical features.

Separation of salts of radium and barium. IVAN YA BASHILOV. Ger. 515,681, July 21, 1925. See Russ. 6040 (C. A. 23, 1059).

Separation of rubidium or cesium salts from alkali salts. KALT FORSCHUNGSANSTALT G. M. B. H. (O. F. Kaseitz and Hans Grasshoff, inventors). Ger. 515,851, Mar. 22, 1930. Complete sepn. of technically pure Rb or Cs salts from alkali salts is effected by pptg. the Rb and Cs as alkali Mg double phosphates. Thus, carnallite contg. 4% RbCl is dissolved in water and treated with $\text{Na}_2\text{H}_2\text{P}_2\text{O}_7 \cdot 12\text{H}_2\text{O}$. The ppt. is Rb-free $\text{MgH}_2\text{P}_2\text{O}_7 \cdot 7\text{H}_2\text{O}$. The filtrate is treated with further $\text{Na}_2\text{H}_2\text{P}_2\text{O}_7 \cdot 12\text{H}_2\text{O}$ and neutralized. The salt $\text{MgRbPO}_4 \cdot 6\text{H}_2\text{O}$ is pptd., free from K. This ppt. is sepd. and dried. A further example is given.

Caustic soda; ammonium chloride. ALFRED MENTZEL. Ger. 510,093, Mar. 27, 1929. NaHCO_3 prepd. by the ammonia-soda process is mixed with C and treated with N at a high temp. The NaCN so obtained is treated with steam, NaOH and NH_3 , being obtained. The latter is returned to the ammonia-soda process and ultimately recovered as NH_4Cl .

Sodium peroxide. HECTOR R. CARVETH (to Rossler & Hasslacher Chemical Co.). U. S. 1,796,241, March 10. Na_2O_2 is made by treating with O a pulverulent mixt. of Na_2O and not over 10% of finely divided metallic Na at a temp. of 200–350°. App. is described.

Titanium dioxide. JOSEPH BLUMENFELD (to Commercial Pigments Corp.). U. S. 1,795,467, March 10. A soln. of a Ti salt such as the sulfate is slowly added to water or to a soln. of lesser concn. while agitating so that colloidal particles of TiO_2 are produced. Cf. C. A. 24, 4175.

Verdigris. A. A. SVITZURN. Russ. appl. 35,471, Nov. 17, 1923. Cu powder, obtained by pulverizing molten Cu, is treated with $\text{Cu}(\text{OAc})_2$.

Treating sulcatea. F. JOURDAN. Brit. 339,023, Oct. 22, 1928. See Fr. 682,704 (C. A. 24, 4504).

Treating greensand and other materials containing potassium. ARTHUR LAMBERT (to Cosmic Arts Inc.). U. S. 1,797,002, March 17. The material is treated with an acid such as HCl or H_2SO_4 and with a basic Na compd. such as Na_2CO_3 to produce a substantially neutral soln. contg. K and Na salts, and the Na is then pptd. as NaHCO_3 , the K salts being left in soln. Various modifications of procedure are described.

Treating kaolin, etc. R. ILNER. Brit. 338,720, Dec. 13, 1928. Materials such as clay, kaolin, diatomaceous earth and sericite earth are improved by treating them in suspension with a slightly sol. salt such as CaSO_4 or CaSO_3 by stirring with such salts (which may be formed *in situ*). The material may also be treated with a bleaching agent or mixed with ultramarine or indanthrene blue for improvement of color. Brit. 338,730 describes stirring the material in acid suspension with $\text{Na}_2\text{S}_2\text{O}_4$, blowing with air, and adding $\text{Ca}(\text{OH})_2$ with or without addn. of blue coloring substances.

Stabilizing solutions of per-compounds. ÖSTERREICHISCHE CHEM. WERKE G. M. B. H. Ger. 518,402, Sept. 22, 1926. See Austrian 119,036 (C. A. 25, 385).

Activated carbon. SOC. DE RECHERCHES ET D'EXPLOITATIONS PÉTROLIÈRES. Fr. 695,212, Aug. 27, 1929. Carbonaceous materials are heated with combustion gases or combustion residues and then activated with activating gases which are of such a compn. that they produce combustible gases. Activating gases which may be used are O, steam or CO_2 , alone or mixed with Cl or HCl . Cf. C. A. 24, 2558.

Reactivating carbon. METALLGES. A. G. Brit. 338,500, June 8, 1928. Reactivation of C in lump form is effected en masse in a container such as an adsorber by preheating at least a portion of the material and then treating it with a gas or gas and vapor mixt. contg. less than 8% O. The preheating is then discontinued and the heating gas supplied at a temp. not substantially greater than 250° and preferably below 200° (thus effecting a localized reactivation zone formation, with travel of this zone in the direction of the gas flow). Cf. C. A. 24, 1188.

Dispersion of carbon black in water. CHARLES R. PARK (to The Goodyear Tire and Rubber Co.). Can. 309,276, Mar. 10, 1931. Carbon black is exposed to the vapors of a material selected from a class comprising petroleum distillate and pine distillate to increase its dispersibility in latex.

Bone black. T. L. WHEELER and J. B. CARPENTER (to Baugh & Sons Co.). Brit. 338,476, April 11, 1929. See Ger. 512,494 (C. A. 25, 1045).

Amorphous carbon from tar. ALEXANDER L. ULLRICH. U. S. 1,796,815, March 17. A solvent such as benzol is added to tar liquefied by heating, and to this mixt. there is added a relatively heavy hydrocarbon material such as fuel oil or kerosene in proportion to cause pptn. of the carbon particles of the mist.; the ppt. is removed and treated with a solvent such as a light oil to remove sol. constituents, and the remaining substantially pure carbon is dried after being sepd. from the assocd. solvent.

Sulfur from sulfide ore. RAYMOND C. BENNER and ALFRED P. THOMPSON (to General Chemical Co.) U. S. 1,795,705, March 10. A mixt. of sulfide ore such as pyrites or pyrrhotite and a carbonaceous reducing agent such as coal is contacted in co-current flow with an oxidizing gas such as air, and the carbonaceous material is employed in sufficient quantity to effect reduction of the S to its elemental form. App. is described. Cf. C. A. 24, 2557.

Condensation products containing sulfur. I. G. FARBENIND. A.-G. Brit. 338,604, Aug. 24, 1929. Condensation products such as those formed from glycerol, ricinoleic acid and phthalic anhydride or similar initial materials, of oily to resin-like character, are heated with S or S compds. such as S monochloride (suitably at 160–70°) to produce products sol. in hydrocarbons and oils and suitable for use in the manuf. of stains, filling or priming compns., cements and varnishes, together with oils, resins, cellulose derivs., softening or plasticizing agents and siccatives. Several examples with details of procedure are given.

Sulfur and bentonite mixture suitable for use in making molded products. CARLETON ELLIS (to Ellis-Foster Co.) U. S. 1,795,364, March 10. A dry mixt. contg. S 90 and bentonite 10% is prepd. by grinding the materials with water and drying the dispersion thus formed at a temp. below the sintering point of S. Quicklime 1% may be added, and the compn. may be used with asphalt for waterproofing roofing, etc.

Artificial substances. SOC. ANON. POUR L'IND. CHIM. A. BALE. Fr. 374,96, Oct. 7, 1929. Addn. to 646,922 (C. A. 23, 2257). Artificial substances are prepd. by condensing in the presence of acid 1 mol. of an aromatic amine with almost 1 mol. of an aldehyde or the equiv. of an agent capable of liberating an aldehyde and converting the fusible resin thus formed, after elimination of acid, either alone or mixed with fillers, dyes, softening agents, etc., to the infusible state with hardening agents. Examples are given. Cf. C. A. 24, 5119, 25, 1046.

Preparing artificial masses. G. S. PETROV and A. K. PETROV. Russ. appl. 35,158, Nov. 12, 1923. Addn. to pat. 360. Artificial masses obtained according to pat. 360 are also prepd. from powdered substances such as are obtained in the treatment of cellulose with hydrating substances, e. g., $ZnCl_2$, $KSCN$ or caustics.

Plastic material. K. I. TARASOV. Russ. appl. 24,198, Feb. 25, 1923, 29,211, June 22, 1928. Addn. to pat. 10,393. Gypsum or a mixt. of gypsum with kaolin or chalk or with hydrocellulose or wood flour is mixed with casein which was preliminarily treated with condensation products of phenol and formalin, and the whole is treated according to pat. 10,393.

Plastic composition. STEFAN DISSMAIER. Austrian 121,553, Oct. 15, 1930. The compn. comprises wood pulp and comminuted leather waste, in approx. equal amts., and a water-insol. org. binder, e. g., resin size. It may be used for coating masonry, protecting steam or water pipes, making buttons, etc.

Adhesive for bronze printing, etc. REINHOLD WOLFFRAM. Ger. 515,522, Mar. 20, 1929. A cement of sufficient adhesiveness for bronze printing, etc., is prepd. by incompletely fermenting a mixt. of dextrin and glucose with yeast and adding H_2O_2 or Na_2O_2 . The fermentation is carried out at about 24°, and 0.3 to 1% NaOH and a small quantity of aq. castor-oil emulsion may be added. A conserving agent such as $BzONa$ may also be added.

Cement for uniting glasses of bi-focal spectacle lenses. A. BANISTER. Brit. 338,555, Aug. 20, 1929. A mixt. is used comprising cellulose acetate, glucose and acetone oil.

"Basing cement" for uniting lamp bulbs to bases. CLYDE C. VAN VOORHIS (to Westinghouse Lamp Co.). U. S. 1,795,746, March 10. A conducting substance such as graphite and an alc.-sol. gum such as shellac are used with $BaSO_4$, marble flour and a volatile solvent such as alc.

Decorating asbestos-cement slabs. HAWENTA-PLATTEN-GESELLSCHAFT. M. B. H. Ger. 477,210, July 16, 1927. See Brit. 337,871 (C. A. 24, 5129).

Wetting, etc., agents. H. TH. BÖRME A.-G. Fr. 37,134, June 18, 1929. Addn. to 671,456 (C. A. 24, 2257). Wetting, cleaning and impregnating agents for use in

the textile and other industries are made by sulfonating acids of high mol wt of the fatty acid or oleic acid series in the presence of anhyd. org. acids, their anhydrides or chlorides. Cf. C. A. 24, 5952

Wetting, etc., agents. I. G. FARBER AND A. G. Fr. 37,153, June 29, 1929. Addn to 632,155. Non aromatic hydrocarbons contg more than 8 C atoms in the mol and their derivs. are sulfonated with gaseous SO_2 , preferably in the presence of an org. diluent. Examples are given.

Wetting, etc., preparations. I. G. FARBER AND A. G. Fr. 37,122, June 24, 1929. Addn to 621,010. Preps. havng the properties of soaps to a very high degree and also a wetting and emulsifying power are composed of sulfonic acids of aliphatic, hydroaromatic, aromatic or heterocyclic compds. having the properties of soaps, or salts of these acids, and, besides, free sulfonic acids or other org. compds. not having in themselves pronounced soap properties. Thus, propylnaphthalene sulfonic acid is mixed with naphthalene sulfonic acid. Cf. C. A. 24, 4900.

Wetting, foaming, emulsifying, etc., agents. I. G. FARBER AND A. G. Ger. 518,408, Jan. 31, 1929. These are prepd. (1) by the reaction of nuclearly halogenated aralkyl halides with aromatic or partly hydrogenated aromatic sulfonic acids during or after the sulfonation process, or (2) by condensing nuclearly halogenated aralkyl halides with aromatic or partly hydrogenated aromatic hydrocarbons and sulfonating the products. Thus, a mixt. of $\text{C}_{12}\text{H}_5\text{I}$ and trichlorobenzyl chloride may be sulfonated, or first treated with ZnCl_2 and then sulfonated. Examples are given.

Detergent. ELECTRIC SMELTING & ALUMINUM CO. Brit. 339,353, Sept. 7, 1929. See U. S. 1,745,844 (C. A. 24, 1712).

Detergents and lubricants containing fatty acid salts. P. SPENCE & SONS, LTD., and T. J. I. CAHILL. Brit. 338,919, May 21, 1929. A fatty acid such as stearic acid is treated with a soln. of alkali combined with a metal oxide sol. in the alkali such as that of Al or Zn, to produce products such as normal or basic double stearates, palmitates, myristates, etc., suitable for thickening lubricants, use with H_2BO_3 in polishing compns., etc.

Polishing agent. ARTHUR VAN DER BRUGGEN. Fr. 37,403, April 15, 1929. Addn. to 667,300 (C. A. 24, 1191). The polishing agent is made by adding the PhNO_2 to the collodion and the acetone to the petroleum to facilitate the mixt. of the petroleum with the collodion in the presence of the PhNO_2 .

Composition for cleaning and polishing silver. CYDERFLA RAISCH. U. S. 1,795,676, March 10. NaCl 5 oz., quinine chloride 1 oz., MgSO_4 1.5 oz. and indigo 4 drops.

Treating magnesium articles to prevent tarnishing. EDWARD C. BURDICK (to Dow Chemical Co.). U. S. 1,795,473, March 10. The articles are subjected to the action of 10-30% H_2SO_4 soln. and then washed.

Bleaching powder. E. KRENS. Brit. 338,940, July 30, 1929. Lime is treated with Cl in a vertical shelf absorber and air is admitted in the bottom portion of the absorber while Cl is admitted at one or more higher points. Various details of the app. and its operation are described.

Trade-marking carbon paper. SAMUEL A. NEIDICH (to Neidich Process Co.). U. S. 1,796,000, March 17. Marking is effected by applying an acid material such as oleic acid capable of increasing the capacity of the web for absorption of waxy, oily transfer material previous to application of the latter.

Carbon papers for manifolding or transfers. P. MEYER. Brit. 338,673, Oct. 11, 1929. Water sol. soap and a water sol. dye are used together (suitably with addn. of a small proportion of oil) in prepg. coating compns.

Transfer material. W. S. LAWRENCE and KAUMAGRAPH CO. Brit. 338,611, Aug. 27, 1929. Transfers are formed with a paper base and a compn. contg ethyl or benzyl cellulose, a resinous material such as cumar, a blown oil such as blown castor oil, and coloring material, and other substances such as Et lactate, tricresyl phosphate, Et acetate and glycol monoethyl ether also may be used in various mixts.

Duplicating printed or written matter. WILHELM RITZERFELD. U. S. 1,795,378, March 10. Reproduction from originals furnished with reverse characters is effected by making the original stencil on a sheet of paper which is thick and soft but has a smooth and dense surface, such as a paper contg chalk, so that upon producing the stencil the characters are impressed deeply into the paper and form raised reversed characters on its back from which copies are made onto previously moistened copy sheets.

Stencil paper. DANIEL A. WILLIAMS (to A. B. Dick Co.). U. S. 1,795,461, March 10. Paper is coated with a compn. comprising a cellulose ester such as nitrocellulose

cellulose acetate, glycerol (suitably in a quantity several times that of the cellulose ether), a relatively small proportion of a resin and a solvent which together form a soft ble product easily displaceable by impact of type Cf C A 24,4598

Sound-record compositions. H J BILLINGS (to A D Little, Inc) Brit 339,335, ly 17, 1929 Records are formed with a backing (suitably formed of sheets of paper eminating with sheets of fibrous material mixed with a thermoplastic binder), a record ect (which may comprise shellac 48, graphite 25 and paper pulp 22% together with C ick and rosin), and a thin coating comprising shellac which may be mixed with a ler and pigment. Various details of manuf are described

Combined sound and picture records. I G FARBENIND A-G Brit 338,817, h. 22, 1929 Sound and picture records are produced on a single film, and both cords are then copied on a single film having a fine grained colored emulsion of steep adation (there being placed before the film bearing the sound record a screen of color mplementary to that of the colored emulsion)

Light-diffusing panels. FRANK COLLINS (to Mutual-Sunset Lamp Mfg Co). S 1,795,294, March 10 A fine-gauge wire screen fabric is immersed in a clear quer which solidifies on exposure to the air and forms a continuous transparent eet encasing the fabric, there is then applied to one surface of this sheet a translucent for medium such as a colored lacquer compn to form a design, and the entire opposite rface is treated with a translucent color medium contrasting in color with the design id providing a background

Freezing-point depressant for use in automobile engine cooling systems, etc. W. W. ORELUP and OSCAR I LEE U S 1,795,854, March 10 Dispersed V oxide used in compns such as CaCl_2 solns in order to protect metal parts of the cooling stem from corrosion. Dispersed compds of Th, Zr, Mo, Ta, Ch, Ti, W or U also ay be used

Dual-effect compression method and apparatus for producing carbon dioxide snow. STUS C. GOOSMANN U S 1,795,772, March 10 Various details of construction and ervation are described

Artificial snow. IGNAZ KREML Austrian 120,806, Aug 15, 1930. Artificial ow for decorative purposes is prepd by pouring hot water on to a mixt. of a solid tty acid and an alkali carbonate and beating the mass

Apparatus (with agitators, screen and brush) for slaking lime. EDWIN J SHAUT o Baker Lime Machinery Co., Ltd) U S 1,796,411, March 17

Substitutes for sugars in industrial processes I G FARBENIND A-G (Otto hmidt and Egon Meyer, inventors) Ger 518,196, June 14, 1927 Water-sol ndensation products obtainable from aliphatic aldehydes are used as substitutes for ucose and other sugars in processes such as *printing or dressing textiles, filling trans- rent soaps, improving the elasticity of cements and adhesives, chrome tanning and gulating latex*. Suitable condensation products may be obtained from CH_3O , glycolic dehyde, glyoxal, or mixts of CH_3O and AcH by treatment with a mild alk. reagent a temp. between 0° and 100° Examples are given

Magnetic structures or cores. JACK C CHASTON and JOHN P. JOHNS (to Western lec. Co) U. S. 1,795,639, March 10 Particles of magnetic material such as a Ni-Fe oy contg. 80% Ni are insulated (suitably by CrO_2 , Na silicate and talc) and the insulated particles are compressed to suitable form and then subjected to the action a fluid such as moist air for a sufficient time to appreciably change the character of e insulating material as by swelling, and the resulting product is then heat treated.

Material for brake linings, clutch facings, etc. AURBORN L HAGER and LON S LAGER. U. S. 1,797,141, March 17 A fabric material is impregnated with a liquid ext. comprising graphite, white lead and creosote U S 1,797,142 relates to manuf f similar material by weaving a fabric and during the weaving impregnating it with a quid mixt. comprising graphite, white lead and creosote, and heating the woven apregnated material to harden the white lead and graphite

Fireproofing peat, etc. E DYCKERHOFF Brit 339,067, Oct 30, 1929 Peat or mlar material is rendered substantially fireproof and non smouldering by treatment ith fireproofing solns such as H_3PO_4 (suitably with heat and pressure), and the treated aterial is pressed and dried and may be formed into peat meal or shaped or pressed rucles

19—GLASS, CLAY PRODUCTS, REFRACTORIES AND ENAMELED METALS

G. F. BARTON, C. H. KERR

Glass K LOENBERG-HOLM *Arch Rec* 63, No 4, 327-58(1930) —The properties of glass as a building material are discussed. I. I. S.

Expansion measurements of several glasses by means of a self-registering apparatus. WILLI M CONY *J Am Ceram Soc* 14, 265-75(1931); cf *C A* 24, 4431 —The app registers photographically. Quartz glass is used as a control material. Slight variations which might escape visual observation can be detected. If results thus obtained are used a close regulation of annealing is made possible. C. H. KERR

Variations caused in the heating curves of glass by heat treatment. A. Q. TOOL AND C. G. EICHLIN *J Am Ceram Soc* 14, 276-308(1931) —Variations in heating curves, due to previous heat treatment, are fully indicative of the nature of such treatment and they also bear a relation to changes in other characteristics of the glass. Differences in heating curves arise from differences in thermal properties, especially exothermic and endothermic effects. It is possible by rapid cooling to suppress the exothermic processes to a great extent and thus prep the way for relatively large exothermic effects on subsequent heating. Annealing procedures may often be manipulated that the resultant changes in certain properties, such as n , will cause a glass to meet certain standard specifications, when otherwise it would not. Also in *Bur Standards J Research* 6, 523-52(1931). C. H. K.

Devitrification F. W. PATSON, *Glass Ind* 12, 1-3(1931) —During the slow cooling down of a soda lime-silica glass melt, the first crystals which form are usually tridymite, wollastonite, or $\text{Na}_2\text{O} \cdot 3\text{CaO} \cdot 6\text{SiO}_2$, especially the last compd., according to Morey, *C A* 24, 5448, rather than cristobalite and wollastonite as stated by Dietzel, *C A* 24, 5447. Diagrams are given showing which cryst. phase may be expected if devitrification does occur in soda lime silica glass of any known compn. I. I. K.

Influence of water content of clays on their properties in drying and firing. V. I. ZUBILANTOV AND D. I. SMIRNOV *Trans Ceram Research Inst (Moscow)* No 24, 30-62(in German 62-5)(1930) —The authors investigated (1) shrinkage and porosity of clays after drying to 110° , (2) shrinkage and porosity of bodies after firing to vitrified states, and (3) shrinkage and loss in wt. at different periods of drying and changes taking place during these processes. Satisfactory results were obtained in measuring the vol. change, loss in wt. and vol. porosity of samples while drying and after firing to Seger Cones 010 and 10. Three periods can be distinguished during drying. In the first, shrinkage and loss of water proceed with const. velocity; shrinkage in vol. of the body remains equal to the vol. of the water evapd. In the second, the decrease in water (pore water) proceeds with changing velocity while the vol. of the clay body remains const. When drying in the air a certain part of the water remains in the clay. During the third period the hygroscopic water is removed by heating to 110° . The quantity of shrinkage water depends on the initial water content, but the quantities of the pore and hygroscopic water do not. The extent of shrinkage depends only on the quantity of the shrinkage water. The relation between the amt. of shrinkage and the amt. of pore water is a characteristic property of each clay. The relative content of pore water in kaolins is larger than in clays, although kaolin is more porous than clays. The water content of the paste has no influence on the change of the d. and size of the clay piece during firing. M. V. KONDOIDY

Ammonium hydroxide as a deflocculant for clay suspensions. ANGELO CASTIGLIONI *Industria chimica* 5, 1465-8(1930) — NH_4OH (0.888), when added in proportions as low as 0.02% by wt. to clay suspensions, increases the fluidity markedly, thus making possible the use of much more concd suspensions. In the manu. of molded ceramic ware it permits the use of suspensions of d. 1.85 which remain as fluid as suspensions of d. 1.72 when Na_2SiO_3 is used as deflocculant. A. W. CONTIERI

Experiments in weathering plastic fireclays. STUART M. PHELPS AND C. G. DENNEY *J Am. Ceram Soc* 14, 319-24(1931) —Changes produced in weathering cannot be readily detected by the usual clay testing methods. Either weathering changes are over-std. or the present test methods are inadequate. C. H. KERR

High-alumina firebricks. J. P. HYSLOP, *Roy Tech Coll Met. Club J.* (Glasgow) 1929-30, No 7, 18-9 —For max. refractoriness, resistance to stresses and resistance to corrosion by slag it is desirable to have a high Al_2O_3 and thus a high mullite ($3\text{Al}_2\text{O}_3 \cdot 2\text{SiO}_2$) content. Spalling tendency is greatly increased by the presence of free SiO_2 . While a high Al_2O_3 content is desirable, refractories with 60% or more

Al_2O_3 tend to become more liable to spalling as the expansion coeff. increases with the Al_2O_3 content. Because of the cost it is economical only in extreme cases to employ materials with high Al_2O_3 derived from diaspor, bauxite, cyanite or corundum. Properly made bricks from clays approaching 45% Al_2O_3 show high resistance to temp., as well as to corrosion and spalling.

H. C. PARISH

Oldenburg clinker. WERNER DOMM *Ziegelwelt* 61, 571-4(1930).—Clinker differs from the ordinary brick in that it has been burned to a sintering point. To make clinker it is essential to use a clay rich in alkali and Fe_2O_3 , but low in $CaCO_3$ and $MgCO_3$. A suitable clay analyzed: SiO_2 70.22, Al_2O_3 13.67, Fe_2O_3 6.80, CaO trace, MgO 1.30, K_2O and Na_2O 3.37, ignition loss 5.30 per cent. The clinker is commonly $22 \times 10.5 \times 5$ cm in size. Well-burned units are very dense, hard, tough and very satisfactory as paving and heavy building materials.

H. F. KRIFGE

Correlation of the crystal structures and hardness of nitrided cases. O. E. HARPER AND GEORGE B. TODD *Fuels & Furnaces* 9, 295-7(1931), *Steel* 88, 51-2. M. S.

Recent developments in the German ceramic industry. FELIX SPRINGER, *Bull. Am. Ceram. Soc.* 10, 85-7(1931).

C. H. KERR

Moisture expansion of glazes and other ceramic finishes. H. G. SCHUECKE AND G. R. POLB *J. Am. Ceram. Soc.* 14, 313-6, *Bur. Standards J. Research* 6, 457-63 (1931), cf. *C. A.* 23, 5554; 24, 3871.—Some glazes show expansion, due to moisture, similar to that of the body. This means less liability to crazing when the ware is exposed to moisture. Artificial weathering treatment (exposure to steam at 150 lb. sq. in. for 1 hr.) caused an av. expansion, in lustrous glazes, of 0.004%, in mat glazes, 0.011%, in vitreous slip finishes, 0.005%, and in porous slip finishes, 0.033%. The moisture expansion of one of the mat glazes was 30% of the av. for a ceramic body with approx. 12% absorption.

C. H. KERR

Electrical porcelain. H. M. KRANER, *Elec. Eng.* 50, 209-11(1931).—American and European samples of standard insulating porcelains show similar characteristics in the finished product despite the variations in raw materials and mfg. processes. A table gives the modulus of rupture of ceramic materials including tests on quenched samples.

W. H. BOYNTON

Comparison of bodies containing blended feldspars and one-mine feldspar of similar composition. C. W. PARMELP AND C. R. ABERG *J. Am. Ceram. Soc.* 14, 309-12 (1931).—Tests were made with an elec. porcelain body mixt. Since it is impossible to consider more than 2 factors in the blending of 2 materials, all but 2 factors should be uniform and this can be obtained only from one-mine materials.

C. H. KERR

Stoneware in the chemical industry. P. WERVAEN *Chemistry and Industry* 50, 213-4(1931).—Because of its strength, its machinability and resistance to acids and bases, stoneware may in the future find more application in the chem. industry. Its high expansion coeff. limits its use. At present the ceramic industry has not succeeded in making a material combining it with the ability to resist high temps. The porosity of stoneware has caused a considerable monetary loss from absorption of precious metals in electrolytic baths. A sillimanite mass can be used for making very dense vessels with an absorption of only 0.01%.

H. C. PARISH

Economy in selection and design of chemical stoneware. J. M. W. CHAMBERLAIN *Chem. Met. Eng.* 38, 142-4(1931).

K. D. JACOB

Manufacture of stoneware. T. W. TALWALKER *J. Indian Ceram. Soc.* 2, 21-27, 65-74, 83-97, 130-43(1930).—Expts. were carried out in a factory mfg. acid jars. Economics of the industry and materials and processes are discussed. T. concludes that it is possible to make a low-temp., highly resistant body, by using frit in the body, and using a fritted glaze high in resistant materials such as ZnO , B_2O_3 and ZrO_2 .

JOHN M. LADINO

The use of refractory natural stone. H. KALPERS *Zentr. Europ. Geisserei-Ztg.* 2, No. 2, 1-3; *Chem. Zentr.* 1929, II, 1196, *Geisserei* 17, 1045-7(1930).—The refractory natural brick of Hugo Schlenkermann (cf. O. Beckmann, *Geisserei-Ztg.* 25, 156-9) contains SiO_2 90.59, Al_2O_3 + TiO_2 6.85, Fe_2O_3 0.20, MgO 0.15, K_2O 0.22 and Na_2O 0.90%. The loss on ignition is 1.10%. The refractoriness of these bricks varies from cone 35 to 38. The behavior of these bricks under pressure was tested with the app. constructed by the Chem. Lab. für Tonindustrie. The bricks were also examd. with regard to their behavior toward changes in temp. and their ability to keep their vol. Their fitness for cupola and reverberatory furnaces is emphasized. Hollow spaces and gaps between the bricks are filled with a special stamped material. G. SCHWACH

Refractory pastes for use in laboratories. MANUEL FELIU, *Quim. e ind.* 7, 280 (1930).—F. mentions various pastes and recommends the Pythagoras paste made by Haldenmayer, porcelain manufacturers, in Spandau, Berlin. It is a mixt. of kaolin,

Al_2O_3 and feldspar. The kaolin decomposes into Al_2O_3 and a glass which with the Al_2O_3 gives sillimanite. Tubes made of this paste withstand all temps. up to 1700° and above. The paste is used in protecting tubes for thermoelec. elements and such tubes are impermeable to gases even above 1700° .
S. L. B. I. HUERTON

Special refractories for metallurgical research. DONALD TURNER *Trans Faraday Soc.* 27, 112-24 (1931).—A description of the refractory materials developed for use in various investigations at the National Phys. Lab., and the methods of production employed.
CURTIS L. WILSON

Investigations on the slagging of refractory materials. IV. Investigations on the corrosive power and the constitution of iron-smelting slags. HERMANN SALMANG and FRIEDRICH SCHUCK *Arch. Eisenhüttenw.* 4, 299-316 (1940), cf. C. A. 22, 3272.—The corrosive power of the oxides present in Fe-smelting slags (CaO , FeO , MgO , MnO , Fe_2O_3 , Mn_2O_3 , Al_2O_3 , TiO_2 , SiO_2 , P_2O_5 and also K_2O , Na_2O , CaS , and CaF_2) on fireclay (at 1410° and 1500°) and on sillimanite (at 1580°) was determined synthetically. Tertiary and quaternary slags being prepared from these oxides. In the fusions, CaO is found to be the strongest oxide. FeO , MnO and MgO following next in strength. Oxides of the type MO (M = metal) react as bases, while those of the type M_2O_3 react as acids. Thus the latter oxides do not attack fireclay and check the attack of bases. In silicate fusions Fe_2O_3 is stable up to 1540° , while Mn_2O_3 decomposes at 1500° . Contrary to expectation from the heats of formation of the silicates of the corresponding oxides, the alkali oxides behave as weak acids at these high temps., while the alk. earth oxides act as strong bases. P_2O_5 checks the attack of bases, and when present in large amts. in the slag, the excess attacks strongly (especially at the higher temps.). The attack on these refractories increases about 50% from 1410° to 1500° and from 1500° to 1580° , from which it is concluded that the viscosity of the melt increases in the same ratio.
J. BALOGHIAN

Preparation of chamotte-free fireproof and acidproof objects. P. P. BUDNIKOV, S. N. ZHEKHARVICH and I. G. SHAKHNOVICH *Ukrainskii Khim. Zhur.* 5, Tech. Pt., 93-103 (in German 104) (1939).—Chamotte-free fireproof and acidproof objects can be prepared by using artificially prepared lean lime to the extent of 85-93% of the mixt. The process lends itself easily to mech. treatment. For the dehydration of lime a drying drum is required working on the counter-current principle. To make brick by this process a pressure of 230 atm. is required. This method eliminates prep. the mixt. before shaping the objects and drying before firing.
S. L. MADORSKY

Standards and tests for coated abrasives. E. C. SCHACHT. *Am. Soc. Mech. Eng.*, Advance Paper, Oct. 16-17, 1930, 7 pp.
E. I. S.

Coated molds (STOCKHOLM) 4. Coating for lamp bulbs (Can. pat. 309,250) 4

Apparatus for feeding mold charges of molten glass. GEORGE E. HOWARD (to Hartford Empire Co.) U. S. 1,796,921, March 17. Structural features.

Device for withdrawing or delivering measured quantities of fused glass. ADOLF SCHILLER. Ger. 519,236, July 29, 1926.

Glass leer of the muffle type with an endless conveyor for carrying articles through it. WM. A. MORTON (to Ansler Morton Co.) U. S. reissue 18,004, March 17. Reissue of original pat. No. 1,684,239 (C. A. 22, 4218).

Glass-gathering apparatus. GEORGE E. HOWARD (to Hartford Empire Co.) U. S. 1,797,200, March 17. Structural features.

Glass furnace and associated apparatus for making spun glass. LOUIS MATHIEU. U. S. 1,796,571, March 17. Structural features.

Apparatus for pouring molten glass into sheets. N.-Y. MAATSCHAPPIJ TOT BEHEER EN EXPLOITATIE VAN OCTROOIEN. Fr. 37,359, Aug. 23, 1929. Addn. to 579,444.

Furnace and associated apparatus for sheet-glass manufacture. SAMUEL C. CRIFT (to Libbey-Owens-Ford Glass Co.) U. S. 1,795,463, March 10. Structural features.

Apparatus for sheet-glass manufacture. JOHN L. DRAKE (to Libbey-Owens-Ford Glass Co.) U. S. 1,796,836, March 17. Structural features.

Apparatus for drawing sheet glass. WERT FRENCH (to Libbey-Owens-Ford Glass Co.) U. S. 1,795,943, March 10. Structural features.

Apparatus for manufacture of sheet glass by rolling. CHANCE BROS. & Co., LTD., and A. L. FORSTER. Brit. 339,535, Jan. 23, 1930. Structural features.

Rolling glass sheets. Y. BRANCATI. Brit. 338,477, April 20, 1929. Various details of procedure are described for insulating glass sheets from their receiving surfaces during rolling, by a gaseous cushion of combustion products.

Apparatus for plate-glass manufacture (roller construction). V. BRANCART. Brit. 339,540, Feb. 14, 1929.

Pressed molded glassware. CHANCE BROS. & Co., LTD., and R. BROWN. Brit. 339,481, Dec. 3, 1929. Various mech. details of manifold are described.

Marbled or multicolored glassware. C. LOCKERIE. Brit. 339,583, April 9, 1929. Masses of glasses are prepared which are of different colors but of substantially the same composition, they are melted in separate pots or furnaces, and one of the glasses is treated with separate quantities of the others. Various details of temp., etc., are given.

Cooling oven for glass. N.-V. MAATSCHAPPIJ TOT REIKER EN EXPLOITATIE VAN OCTROOIE. Ger. 518,193, June 15, 1929. Addn. to 478,927 (C. A. 23, 5785).

Blowing and molding articles of fused silica. QUARTZ ET SILICE. Brit. 338,456-7, April 16, and 17, 1929. Mech. features.

Blown articles of fused silica. QUARTZ ET SILICE. Brit. 339,107, Oct. 30, 1929. Various details of manifold are described.

Fusing together grooved sections of fused quartz, etc. LEON SAMUEL (to Thermal Syndicate Ltd.). U. S. 1,700,401, March 17. Mech. features.

Uniting glass sheets with non-brittle transparent material. LIBBY'S OWENS GLASS CO. Brit. 338,988, Aug. 30, 1929. The surface of the intermediate sheet is roughened by spraying onto it a discontinuous layer of an adhesive material such as a mixt. of equal parts of an ester of phthalic acid such as dibutyl phthalate and of a polyglycol deriv. such as monoethyl ether or diethyl glycol. The inner surfaces of the glass sheets to be coated with a soln. of a cellulose deriv. and a synthetic resin, with or without a plasticizer, and may also be coated with adhesive and the sheets are united under pressure while heated. Cf. C. A. 25, 1942.

Reinforced glass. SOC. DES PROPRIETAIRES CHIMIQUES RUOYER PORCELAINE. Fr. 57,119, June 22, 1929. Addn. to 188,563 (C. A. 23, 5291). The compound sheets of glass are heated without pressure to a temp. above the ordinary after removal from the bath of plasticizing material.

Bricks. FRANCO RANDINI. Austrian 120,972, Aug. 15, 1920. Clay congl. at least 20% of water is warmed to about 95° and molded in a press maintained at the same temp. The bricks are dried at once in a moist atm. in an oven in which the temp. is initially the same as that of the bricks, but is gradually raised.

Bricks and mortar from magnesite or magnesite. MAANSHITAPAK, LTD. Hong. 101,355, July 20, 1929. Overheated natural or artificial magnesite is powdered, particles smaller than 0.1 mm. are removed, and 5 to 35% of a substance rich in Al matter, e. g., bauxite, corundum or corundum slag, is added. Bricks made of this mixt. are very refractory.

Tunnel kiln for bricks. VOJTECH LAZROVICKA. Austrian 120,977, Aug. 15, 1920.

Extracting clay and other aluminous raw materials with acids. OSKAR JONAS, KURT WIGER, and GOTTHARD TREBITZ (to I. G. Farbenindustrie A.-G.). U. S. 1,700,107, March 10. Material such as calcined clay is treated with acid such as HCl in large batches *in situ*. The reaction is started by flooding the charge with acid washing water from a previous operation and adding in preheated condition a small portion of the total acid required for the reaction, with further acid added to obtain a strong alumina soln., simultaneous withdrawal of said ext. to maintain a const. level of liquid on the charge and maintenance of a min. temp. of 105° in the reaction zone by the heat of reaction.

Ceramic noble metal preparations. DEUTSCHE GOLDSILBER-SCHNITTSTANFALFABRIKAL ROSSLER (Franz Schellung, inventor). Ger. 515,588, Oct. 14, 1927. Noble metal preps. for decorating and coloring ceramic ware are obtained from dispersions of the metal or its compds. in cellulose derivs., in the presence or absence of solvent and flux. Thus, a 30% AuCl₃ soln. in alc. is stirred into a soln. of nitrocellulose in ethyl glycol and cyclohexanone. Several addnl. examples are given. The method is stated to be applicable to all metals.

Apparatus for dipping pottery or earthenware into glazing or coloring liquid. G. WADE & SON, LTD., and G. A. WADE. Brit. 339,303, Sept. 27, 1929. Structural features.

20—CEMENT AND OTHER BUILDING MATERIALS

J. C. WITT

A new shaft-kiln cement plant. HORST LAEGER. *Techn.-Ztg.* 55, 322-5 (1931).—The shaft kiln with rotating grate recently erected in Ankara (Angora) is described.

F. O. ANDERSON

The enumeration of French cement plants. PAUL RAZOUS *Ciment* 36, 13-8, 50-61 (1931) — A list is given of the different companies and plants in France making the following cements: natural, Roman, artificial (portland), high-early-strength, alumina, slag (with lime), grappier (from flint nodules), iron (70% portland with 30% granulated blast furnace slag) and blast furnace (containing less clinker than iron cement)

F. O. ANDEREGG

Modern cement burning. HENNAICH LUTTSCHITZ. *Tonind.-Ztg.* 55, 184-6 (1931) — A description is given of the Schröder kiln, in which the dry, powdered raw material falls through a shaft heated by powdered coal falling through an outer annular space (Ger. pat. 495,001, C. A. 24, 3310). The Rigby patent (Ger. pat. 478,632, C. A. 23, 4702) for spraying slurry into a kiln is also described

F. O. ANDEREGG

Cement synthesis in the four-component parallelogram and the degree of lime saturation according to Köhl. M. SPINDEL. *Tonind.-Ztg.* 55, 305-7 (1931) — The parallelogram for the 4 components CaO , SiO_2 , Al_2O_3 , and Fe_2O_3 may also be conveniently used for Köhl's degree of lime saturation. HANS KETTEL. *Ibid.* 320-2 — Köhl prefers his quotient based on mol. wts. M. SPINDEL, HANS KETTEL. *Ibid.* 360-8 — More polemics

F. O. ANDEREGG

To secure cement of uniform quality. W. SACHS. *Tonind.-Ztg.* 55, 291-3 (1931) — Proper feeding, weighing and mixing, obtained with suitable automatic equipment, are useful in maintaining a uniform quality

F. O. ANDEREGG

High-early-strength cements and their increases in strength. P. MAY. *Tonind.-Ztg.* 55, 171-3 (1931) — Generally increasing strengths have been encountered up to 5 years, with occasional temporary retrogressions

F. O. ANDEREGG

German specifications for alumina cement. HUGO VIEBALLE. *Tonind.-Ztg.* 55, 325-6 (1931) — Fineness is to be not more than 2% on the 75- and not more than 15% on the 180-mesh sieve. Tests are to be made at 15° to 20°. Specimens are to be stored 24 hrs. in a damp closet and then under wet cloths. The min. strengths are to be compression 400, 500, 600 and tensile 25, 30 and 45 kg per sq. cm. after 24 hrs., 28 days' wet and 28 days' combined storage, resp.

F. O. ANDEREGG

Tentative specification for high-alumina cements. *Structural Eng.* 8, No. 10, 369-73 (1930)

E. I. S.

Asbestos-cement pipes—their production, characteristics and uses. A. SCROTTAK. *Gas u. Wasserfach* 74, 293-7 (1931) — Up to the present, the use of asbestos-cement pipes has been limited because of their small production. These pipes are made by mixing 1 part of ground asbestos with 5 parts of cement and sufficient water to give a thin grout. This grout is then collected in a 0.2-mm. layer on an endless belt which transfers the material to an iron roll of the internal diam. of the desired pipe. This roll presses out the surplus water, and when the coating becomes heavy enough the roll is removed and a new roll substituted. The tube is removed from the iron roll while still soft and placed on a wooden mandrel of similar size. The pipe is then allowed to harden in a damp, fairly warm room. These pipes have found use for flue pipes, house-drain lines, etc. Some difficulties have been encountered with their use for gas and water, but these difficulties are being solved and the pipe can be used up to moderate water pressures. The pipes can be sawed, burned or bored with ordinary wood working tools. They may be dropped 6 ft. on a concrete floor without damage. Breaking strengths, etc., are given, as well as hydraulic tests. Joints are described and illustrated. The pipe is said to cost about $\frac{1}{2}$ less than cast iron or steel pipe

R. W. RYAN

Rapid determination of magnesia in portland cement. CH. ROCOZINSKI. *Tonind.-Ztg.* 55, 355-6 (1931) — Remove silica and add $N\text{NaOH}$ till a ppt. forms and the supernatant liquor is clear. Filter into an excess of 0.2 $N\text{HCl}$, and titrate the excess with 0.2 $N\text{NaOH}$ with phenolphthalein indicator

F. O. ANDEREGG

Zeolites and hydraulic cements. HENRI LAFUMA. *Rev. matériaux construction trav. publics* 1931, 1-5, 45-9 — The properties of hydrated Ca aluminates and silicates are reviewed and compared with those of zeolites. The best explanation of the setting and hardening of the silicates seems to be the development of a network structure

F. O. ANDEREGG

The influence of free lime in cement clinker on the solubility of lime and clay. KATSUO KOYANAGI. *J. Soc. Chem. Ind., Japan* 34, Suppl. binding 56-8 (1931), cf. C. A. 25, 1353 — Well burned, low lime clinker shows high soly. for clay and low for lime. An increase in lime greatly decreases the soly. of clay and increases that of lime. Low lime cement when mixed with H_2O forms K aluminat crystals first, then $\text{Ca}(\text{OH})_2$, and gels; with high lime the process is reversed.

V. F. HARRINGTON

Suitability of argillaceous sands for cement mortar and concrete. WERKEN. *Zentr. Bauverwaltung* 50, 671-3 (1930) — Report on original tests, using sands with per-

centage of impurities up to 13.6 W. advises strongly against the use of impure sands. E. I. S.

Determining the proportions of mortars and concretes. LUDWIG SCHUSTER. *Tonind Ztg* 55, 224-5, 243-4(1931).—Examples are given of analyses of old mortars and concretes from which the approx. original compn. can be calculated. The analysis of the aggregate, if available, simplifies the calculation. F. O. ANDEREGG.

Theory of composition and structure of concrete. A. M. GÖNBERG. *Rev. matériaux construction trav. publics* 1930, 477-81, 1931, 21.5.—The space between the largest aggregate grains in concrete may range from 73 to 29%. To obtain the latter, very exact placing of each particle is required, which is hard to accomplish. Therefore, an intermediate range is selected with the ratios of sizes lying between 1:15:225 and 1:30:900, e. g., cement 0.03 mm., sand 0.45 to 0.9 mm., gravel 7 to 27 mm. and plumb 100 to 800 mm. This type of grading gives good packing. The water content is very important. G. suggests that the addition of solutes, by lowering surface tensions, should reduce the voids in concrete. To get a porous concrete select those sizes and shapes which give poorest packing. F. O. ANDEREGG.

The uniformity of concrete and its control on the job. FRITZ EMPERGER. *Tonind Ztg* 55, 382-3(1931).—Cubes and beams are made on the job and broken at 7 days. F. O. ANDEREGG.

Calculation of the best composition of aggregate for concrete. H. RICHARZ. *Tonind Ztg* 55, 113-4(1931).—An example is given of a suitable blending of Rhine sand and Monier gravel that produced a grading closely approaching the ideal. F. O. A.

Testing protective coatings for concrete. P. MUCKE. *Tonind Ztg* 55, 322(1931).—For accurate solvent detn. in bituminous prepns., place 5 g. in a small foil box containing a small roll of gauze. After one hour at 100° distribute the bitumen evenly over the gauze and sides of the box and dry to const. wt. at 110°. To test the resistance to chem. effects cover a mortar slab with three coats after a cylinder has been cemented to the slab with neat cement. Place the soln. to be tested in the cylinder. F. O. ANDEREGG.

Cracks in concrete. M. E. ROSSNACH. *Rev. matériaux construction trav. publics* 1931, 12-7.—Cracks developing in concrete result from a combination of shrinkage and insufficient tensile strength. Emphasis should be laid on securing a greater tensile elasticity in concrete, even at the expense of compressive strength, because concrete seldom fails in compression. F. O. ANDEREGG.

Researches on the protection of concrete against corrosive waters. OTTO GRAY. *Zement* 19, 936-41, 970-4, 995-8, 1041-3, 1066-8(1930).—The study of the corrosive action of water on concrete involves many factors beside the concn. of salt solns. Corrosive action increased with stirring of the solns., coarseness of cement used in the test pieces and poor gradation of cement particles, fineness of the sand used and quantity of mixing H₂O used, a deficiency being as objectional as an excess. The addition of 3% clay to the specimens decreased their resistance to MgCl₂ solns., if they had not been air-cured. The compressive and tensile strengths were not decreased by clay addition. A curing period of 21 days in air after 6 days under H₂O made the mortar specimens (1:3) resistant to the attack of a 10% MgCl₂ soln. during 20 months' exposure. H. F. K.

The extraction of bitumen for testing from street paving material and the determination of sulfur in bitumen. FR. SEELIG. *Chem.-Ztg* 55, 145-6(1931).—Directions are given for the extrn. of the bitumen (the method is claimed to be more accurate and reproducible) and for the complete combustion of the bitumen in the detn. of S. V. F. H.

Moldering of natural building stones and prevention of their quick decaying. O. LÁCZKOS-FRITZ. *Magyar Mérnök-Egylet Évkönyve Kélműve 7, 1-21(1930)*.—The process of moldering is described. The ideal preventive agent should keep water out but should let the original moisture evaporate from the stone. The stones of the Hungarian Parliament Building were treated by Möller's method. This consists in spraying the stone surface with a paraffin liquid and repeating this process twice, each time after complete drying. The impregnation requires warmth, so is easiest in summer. If the stone is not warmed enough by the sun's rays, it must be heated artificially. The surface should be quite pure and dry. The cost of the treatment tried at the Cathedral at Pécs did not reach 40 (Hungarian) cents per sq. m. of surface. S. S. DE FÉALY.

Concrete coverings for pipe lines (HUGH) 22. Separation and size distribution of microscopic particles (ROLLER) 2. Substitutes for sugars [in improving elasticity of cements and adhesives] (Ger. pat. 518,196) 18. Compositions containing rubber and cement (Brit. pat. 339,002) 30. S and bentonite mixture for use [as roofing] (U. S. pat. 1,795,364) 18.

is well mixed with about an equal amt. of water before the combustible material is added. A small quantity of HCl may be added to the plastic mass before firing.

Artificial marble. BARTHOLMEY J. BAUNEL. Fr. 37,177, Jan 6, 1929. Aided. to 597,896. Cerrotin minium is used with MgCl_2 and MgSiO_3 for the production of artificial marble.

Making artificial marble from cements of various colors. KARL WALDFERT. Austrian 121,563, Oct. 15, 1930. Manipulative details are described.

Using the residues of the artificial marble industry. ALEXANDER STAFGE. Ger. 515,519, Apr. 16, 1930. Aided. to 511,710 (C. A. 25, 2364). The MgO content residue from the artificial stone and marble processes is extracted with dil. HCl and the remainder used as flux for portland cement.

Blocks and other shaped articles of alabaster and gypsum. F. J. G. GARNER. Brit. 338,165, Feb. 28, 1930. In order to give the articles the hardness and appearance of marble they are dried by heating to 150–200°, impregnated with a solution formed of water 1 gal., cone. Na silicate solution 1 lb., MgSO_4 1 oz. and fused CaCl_2 1 oz., and then polished and rubbed with raw linseed oil.

Roofing material. GEORGE D. CRANES. U. S. 1,796,861, March 17. A flexible foundation material containing a major proportion of mineral fiber such as asbestos is saturated with a relatively hard asphalt having a penetration not exceeding 11 at a temp. of 25°.

Roofing composition. GEORGE A. OSTERDAY (one-half to Edward Dome). U. S. 1,799,474, March 17. A mixt. is formed of pitch 25–40, asphalt 25–50, cork 25–50 and asbestos 5–10%.

Dampproofing and waterproofing walls. JOSEPH ROSE. U. S. 1,799,269, March 10. The inner surfaces of walls are coated successively with a cementitious material such as a portland cement mixt. containing emulsified asphalt, an overlying bonded coating of waterproofing material such as an asphalt paint and a finishing coating such as plaster, which is protected from dampness seeping through the wall by the underlying coatings.

Wires and wire netting for carrying plaster. C. SCHADDER. Brit. 338,755, Dec. 10, 1928. The wires are first coated with an adhesive such as glue, then treated with sand or volcanic ashes to form a facing, and then sprayed or washed with an aqueous concrete mixt. or the like (leaving the meshes of the netting open).

Drying lumber, etc. HENRY W. COWAN. U. S. 1,799,141, March 10. The material is placed in a chamber the air in which is heated to a temp. capable of damaging the material on prolonged exposure, and hot moisture laden air is periodically removed while dry air simultaneously admitted is to cause periodical reductions in temp. to below the damaging point through absorption of moisture from the material by the newly admitted dry air (which is heated to a suitable temp. in the chamber). App. is described.

Apparatus for partly impregnating timber. FRANZ KRAUSZ and LUDWIG TRAMER. Austrian 121,256, Sept. 15, 1930. Means is described for sealing the tank in which the trunk is partly immersed.

Impregnating wooden loom shuttles. I. G. FARRERINO A. C. Brit. 338,976, Aug. 30, 1929. The wood is impregnated with chlorinated organic compounds such as chlorinated naphthalenes or solid chlorinated biphenyl, suitably with conjoint use of a substance of lower m. p. such as paraffin or montan wax. Various details and examples are given.

Preserving wood. SOC. DE RECHERCHES ET DE PERFECTIONNEMENTS INDUSTRIELS. Brit. 339,217, April 8, 1929. The process of wood impregnation described in Brit. 310,805 (C. A. 24, 707) is modified by the use of a solution of As sulfide in an alkali sulfide such as Na sulfide as the first injected solution which may also contain acetates and phenols, and this treatment may be followed by a second impregnation as described in Brit. 310,804 (C. A. 24, 707). Various details and modifications of procedure are described.

Wood-preserving composition. BASILIUS R. V. MALENKOVIC. U. S. 1,795,658, March 10. A stable crystalline product of 2,4-dinitro-1-chlorobenzene and α -naphthylamine is used, suitably with various other preservatives.

21—FUELS, GAS, TAR AND COKE

A. C. FIELDNER AND ALDEN H. EMERY

Alcohol automobile fuels. K. R. DIETRICH. Chem.-Ztg. 55, 215–6 (1931).—See C. A. 25, 576

The use of alcohol for motor-fuel mixtures. KARL LOSKOT. Chem. Listy 25,

nil in the first 2 ranges to 755 and 2165 cu It per ton, resp., in the last 2 temp ranges. With immature brown coals it is always possible to distinguish 3 successive and independent changes, resulting in successive expulsions from the coal substance of (1) CO_2 and H_2O , as if by internal condensation, (2) methane, higher paraffins and olefins, as if by the elimination of aliphatic side chains from the coal complex, oils also appearing, and (3) H. With increasing maturity of coals these successive steps overlap and are not easily distinguishable on distg well-matured bituminous coals. The "Sorblet method" of benzene-pressure extrn. of coals at 250, 500 and 700 lb per sq in and results obtained thereby are outlined. The 4 different fractions obtained by this method from bituminous coals are described and their effects on coking properties discussed. Recent results from extrn. expts appear to furnish an explanation to the problem of the development of coking propensities during the maturing of coals. Extrn. of certain immature brown coals yielded none of the usual Fractions III and IV but in their place a quantity of phenolic compds (among which phenol, *p*-cresol and catechol were identified) and phenolic esters. Expts on Canadian lignitic coals of somewhat greater maturity showed that brown lignites from Saskatchewan yielded phenolic compds instead of fractions III and IV while with laminated black lignites from Alberta phenolic compds were not obtained, but in their place a substance simulating fractions III and IV began to appear. These laminated black lignites appear to represent a transition stage in the development of fractions III and IV obtained from bituminous coal with coking properties. This conclusion is strengthened by recent exptl. proofs of the benzenoid character of fraction IV yielded by coking coals. The yields of fractions I, II, III and IV totaling 1.25–15.6% of the coal substance obtained from 9 different kinds of coal are tabulated. The last section of the paper is on the benzenoid character of the main coal substance (85% + of the coal) as shown by its oxidation with alk. permanganate (C. A. 24, 5967). The C/H ratio of the original dry coals and of the dry residue after benzene pressure extrn. and the percentage of volatiles of the original dry ashless coal and of the dry ashless residues are given for 6 coals varying from immature brown coal to strongly coking bituminous coal. The C/H ratio was always higher in the dry residues and the percentage of volatiles always less than in the original coal. The wts of acetic, oxalic and crude benzenoid acids obtained from 3 of the coals and the C/H in crude benzenoid acids are tabulated. A table on C balance of the oxidation shows the percentage of original C of the coal substances appearing as CO_2 , acetic, oxalic, succinic and benzenoid acids. With the Morwell brown coal total C accounted for equaled 99.7%, with the Durham "Busty" coking coal 99.4%. The oxidation products from 3 of the coals have been completely investigated. From the Canadian Estevan brown lignite the products were: CO_2 and the acids: acetic, oxalic, succinic, phthalic, isophthalic, terephthalic, trimellitic, bemimellitic, trimelic, pyromellitic, mellophanic, benzene-pentacarboxylic and mellitic. The same products except trimelic acid were obtained from Australian Morwell brown coal and with the exceptions of succinic, bemimellitic and trimelic acids from Durham Busty coking coal. The possible benzene ring condensation compds or condensed nuclei with aliphatic side chains present in the coal substance which give on oxidation the various acids isolated are discussed. The large new fields in coal chemistry opened up by this discovery and the possible great economic value of continued research along these lines are suggested. An extended discussion follows the article.

W. W. HODGE

Proximate analyses of coals from North China. S. D. WILSON. *J. Assoc. Chinese and Am. Eng.* 11, No. 7, 9–22 (1930), cf. C. A. 24, 5169. E. I. S.

Unit coal studies on some Virginia coals. F. H. FISH AND J. A. ADDLESTONE. *Ind. Eng. Chem., Anal. Ed.* 3, 155–8 (1931).—The 3 formulas most commonly used for correcting the calorific value of coal to an ash free basis are given and briefly discussed. To det. which formula could best be used in correcting ash to obtain unit B. t. u. for some Virginia coals, expts. were carried out on coal from 10 different seams and on 28 samples from the Merrimac seam. Coals from the 10 seams were ground to 20 mesh and float-and-sink sepsms made with CCl_4 + C_6H_6 of sp. gr. 1.35. Two tables give: (1) percentages of ash, S, observed B. t. u., calcd B. t. u. by the 3 formulas and differences for untreated, float and sink coal from the 10 seams, and (2) percentages of ash, S, observed B. t. u. and calcd B. t. u. for the 28 samples of Merrimac coal with deviations from the av. B. t. u. From these data the unit B. t. u. values calcd by the Fieldner modification of the Parr formula were in best agreement. However, the differences were rather large and a second series of expts. was run to correct for CO_2 . A description and a diagram of the app. used for detg. the CO_2 in coal are given. Two other tabulations summarize the new data obtained by modifying Parr's formula to correct for CO_2 : unit B. t. u. = $(\text{B. t. u.} - 5000 \text{ S}) / (100 - [0.08(\text{ash} + \text{CO}_2) + 22/40\text{S}])$ and

Fieldner's formula unit B t u = $(B t u - 4650S)/(100 - [108(ash + CO_2) + 7/5])$. With these modified formulas the unit coal B t u's are in closer agreement. For the 10 coals examined the av. of differences in unit B t u between the untreated, float and sink portions of each coal was 52 B t u by the Parr formula and 96 B t u by the Fieldner formula. For the 28 samples from the same seam the av. deviation from the mean is 75 B t u by the Parr formula and 86 B t u by the Fieldner modification of this formula. These variations may be due to the inherent variations in the different layers of the coal seam but if this factor is neglected the av. deviation is not much greater than the exp'd error in the B t u detn. W. W. HODGE

Hungarian coal and coal-oil problems. I. VÉRTÉK. *Bányás Kohás Lapok* 62, 299-314 (21 32 117 39)(1924).—Expts. were carried out with highly bituminous brown coals of the Upper Oligocene of Bakony Mountain in Hungary. The quality of the ext'd bitumen and of low temp. tars corresponds with that of pyropisite found near Halle, Germany. S. S. DE LÉVELY

Preparation and drying of briquetting coals. A. GEBHAARDT. *Braunkohle* 28, 282-304(1929). E. I. S.

Temperatures in briquets during briquetting. H. SCHUSTER. *Braunkohle* 28, 185-22(1929).—Three methods of detg. temp. of 7-in. room heating briquets during briquetting are described. E. I. S.

Ash and clinker of upper Silesian coals in furnace operation. PAUL FUCHS. *Arch. Wärmewiss.* 12, 7-9(1931).—The compns. of the coarse stone and the finely divided ash in the coal proper are very different. The fly ash resembles the latter more than the former. It is not possible to tell from the compn. of a refractory whether it will be attacked by the ash. Effects of local overheating, both on brickwork and on iron parts, are often wrongly attributed to the slagging effect of the ash. L. W. TIEBLE

Portable equipment for crushing and quartering samples of coal, coke or other lumpy materials. L. S. PERRYJOHN. *Ind. Eng. Chem., Anal. Ed.* 3, 163-4(1931).—Dimensioned drawings are given with the descriptions of a portable steel quartering hopper and crusher. The hopper is an inverted cone and the coal is discharged from it on to crossed angle irons and received into 4 triangular boxes of approx. 1 1/2 cu. ft. capacity. The crusher consists of a steel plate perforated with 1 in. holes, surrounded by an 18" high hopper and with a receiving drawer fitted below the plate. A steel tamper is used to crush the coal until it passes the holes in the plate. Directions for using the app. are given. P. has used the app. for sampling coal up to 8 in. lumps. W. W. HODGE

Short tube mills for pulverizing coal. C. PYRAM. *Rev. matériaux construction trav. publics* 1931, 23-9.—For efficiency and low cost of installation, short tube mills with circulating loads have many advantages. F. O. ANDEREGG

The present state of the pneumatic process for dry-cleaning coal. L. H. III. KENEL. C. APPLEBY AND EDWARD O'TOOLE. *Blast Furnace and Steel Plant* 18, 1834-5, 19, 268-72, 418-22(1931). E. H.

Industrial heating with powdered coal in the inorganic industries. J. DEFORGE. *Rev. matériaux construction trav. publics* 1930, 469-73, 1931, 8-10, 49-53, 101-5.—The advantages include complete combustion of the fuel with ready starting, stopping and control, and the possibility of using lower grade fuels. The disadvantages are danger of accidents, which can be largely overcome by proper design and training of personnel, the difficulty of handling cinders, which can usually be handled in designing the equipment, and the cost of installation, which is sufficient to require a very careful study before making a new installation. F. O. ANDEREGG

Burners for pulverized-coal boilers in their relation to development of combustion chambers. J. GROSZ. *Braunkohle* 28, 389-95, 413-9(1929). E. I. S.

A useful combustion chart. A. ALISON. *Colliery Eng.* 8, 102-4(1931).—A chart is described for detg. excess air in the combustion of coal, depending on the proportion of N and uncombined O present in the flue gas. ANN NICHOLSON HIRN

Valuation of coal used for heat source. H. MATSUNAMI. *Japan Geol. Res. Bull.* (Tokyo) 18, No. 34, 25 pp.—The meaning of the chem. constituents of coal and their influence on combustion are discussed. E. I. S.

Lignite research and pollen analysis. F. KIRCHHEIMER. *Braunkohle* 29, 418-63 (1930).—The microflora, especially pollen analytical, method as an aid in lignite research is discussed. Bibliography. E. I. S.

Improvement of Hungarian lignites by drying. G. SZELÉNYI. *Magyar Műnök Értékeztető Közönye* 63, 215-32(1929).—Drying can be accomplished with good results either by Fleissner's method (treating with steam at 10-20 atm.) or by the method adapted by Mátravideki Coal Mine, Ltd. (boiling with high-pressure water). The

removal of 1 kg moisture requires 0.6 kg steam. Dried lignite, especially that dried by the 2nd method, has good grain size and contains little waste dust. The ash content diminishes about equally in both methods. Reply. H KLEIN *Tűzeléstech* 2, 39-42(1930)—K. states that Fleissner's method is more useful for effecting drying. Reply. G SZELÉVYI. *Ibid* 43—S insists on his former views on the basis of his own expts. S S DE FINÁLY

Preparation of pulverized lignite with pneumatic circulation drying. P. ROSIN AND E RAMMLER. *Braunkohle* 29, 557-63(1930)—Circulation drying is a pneumatic process according to which the transport and drying of wet products in suspension are effected by hot gases which circulate at high speed through a pipe system. Details of a circulating drier are given. E I S

Calculations and investigations of lignite-briquet cooling. H SCHUSTER. *Braunkohle* 28, 305-18(1929)—It is technically possible to cool briquets to about 2 to 3% of the surrounding temp. and loading at this temp. without the possibility of further cooling in car is considered safe in respect to quality of briquet and self ignition hazard. E I S

By-product ammonia. H HOLLINGS AND E W SMITH. *Inst. Gas Eng. Communication* No. 19, 25-30(1930)—A discussion is given of possible ways and means for increasing the net value of NH_3 at the gas works, which in many cases under existing conditions does not pay the cost of its removal from the gas. Production in Great Britain is equiv. to 150,000 tons of $(\text{NH}_4)_2\text{SO}_4$ annually. At the present time the Comm. cannot indicate any new research in this field with any reasonable prospect of economic advantage. H L OLIN

Preparation of ammonium sulfate in the form of coarse thick plates. W GLOUW, V KLEMPF AND H RITTER. *Ber. ges. Kohlentech.* 3, 371-84(1931)—The effects of various conditions upon the cryst. form of coal-tar $(\text{NH}_4)_2\text{SO}_4$ were studied. Temp. and rate of crystn. and the quantity of such impurities as pyridine, tar acids, phenol and As appeared to have no effect upon the form of the crystals, but Fe did appear to be the detg. factor. To det. the effect of Fe salts, the Fe present in the crude liquors was pptd. out with H_2S and then definite quantities of Fe salts were added. Fe^{2+} ion had an injurious effect upon the cryst. form, only a fine crystal brew being obtained in its presence. Cr and Al have the same effect, but the lower-valence forms of Fe, Mn and Co upon the other hand have very favorable effects. The quantities of Fe, Mn and Co used were 2.4 g. of sulfate per l. If it is desired to get the same cryst. form of sulfate as is obtained from synthetic NH_3 , it is necessary to remove the Fe. This is accomplished by oxidation to Fe^{3+} and pptn. as hydroxide by making alk. with NH_3 . H STORZT

Naphthalene removal with tetralin. P. DEURSCHE. *Gas u. Wasserfach* 74, 245-7(1931)—A satisfactory naphthalene solvent to be introduced into gas mains must be (1) a good solvent for naphthalene, (2) of low vapor pressure to avoid too great losses, (3) of higher vapor pressure than naphthalene in order to be present in sufficient amts. to dissolve the condensed naphthalene. The vapor pressure of naphthalene and a series of possible naphthalene solvents, as well as the soly. of naphthalene in these solvents, are given for 0° and 20°. Tetralin satisfies the above conditions best. A given gas contained 13 g. naphthalene per 100 cu. m. On cooling from 20° to 0° 8.5 g. of naphthalene will be deposited per 100 cu. m. of gas. 62.5 g. of tetralin is required to dissolve this amt. of naphthalene at 0° (satd. soln.) and 60 g. to sat. the gas at 0° or a total of 122.5 g. of tetralin per 100 cu. m. of gas. The vapor pressure of tetralin is such that gas at 20° can carry 200 g. per 100 cu. m. Xylene can dissolve 20.8 g. of naphthalene per 100 g. at 0° so only 32.3 g. of xylene is required to dissolve the naphthalene. However, the gas requires an addnl. amt. of 1690 g. xylene to sat. itself at 0°, or a total of 1722 g. must be added to the gas to insure sufficient xylene condensate to dissolve the naphthalene when the gas is cooled to 0°, over 10 times as much as with tetralin. At Bremerhaven tetralin costs 55.25 marks per 100 kg. The total cost of adding tetralin for the first year was 2.9 pf. per 100 cu. m. The typical app. for vaporizing tetralin is described and illustrated. Experiences in other cities are given. D regards tetralin as the single satisfactory means for combating naphthalene and removing tarry deposits from mains. R. W. RYAN

Esthonian combustible shales. OTAKAR VONDRÁČEK. *Chem. Listy* 24, 471-3(1930)—Combustible shales are located between the cities Tapa and Narva and cover 2470 sq. km. The shales contain 70% org. matter which came from a rich sea fauna; the shale burns with a luminous flame and emanates an oily odor. The shales are used in the industry and in firing locomotives, a great deal of fine ash forms which penetrates everywhere. The ash nuisance is obviated by using an oil prep. by the destructive distn. of shale. FRANK MARESH

The Italian gas industry. R. DECKER. *Gas u. Wasserfach* 74, 293-4 (1931) —
 (1) 11 lights are given for the Italian gas industry. R. W. RYAN

Determination of the heating value of gases by explosion. HANS LÖFFLER. *Arch.
 chem. u. phys.* 11, 4 (1931) et C. A. 25, 237. An app. is described and illustrated in
 detail, moved with a definite vol. of air and exploded, the heating value being
 determined in temp. of a Hg thermometer placed in the center of the pipet, which is
 filled with gas. The calibration appears to be empirical. R. W. T.

The 25th report of the Joint Research Committee of the Institution and Leeds
 University. Examination of the products of combustion from typical gas appliances. V.
 THOMAS HARDIN. *Inst. Gas Eng. Communication*, No. 16, 60 pp (1930) — Tests on
 properly adjusted surface combustion water heaters showed CO in waste gas less than
 0.1%. Similar tests with inverted incandescent mantle burners indicated a higher
 rate of CO production than with the Bunsen flame especially at gas rates lower than
 normal. The rapid increase of CO when the burner was turned low points to the ad-
 vantage of using a smaller unit at full capacity when little light is required.

H. L. OLIN
 Some further experiments on the combustion of inflammable gases by electric
 sparks. J. D. MORAN. *Phil. Mag.* [7] 11, 358-63 (1931), et *Phil. Mag.* 43, 259
 (1923). C. A. 17, 2505. — Weak mixtures of coal gas with air at atmospheric pressure
 and pressure were subjected to a stream of high tension sparks, maintained without burning
 in the gas. Combustion in the mixture is proportional to the heat energy of the spark
 discharge when the gap width is constant, and proportional to the gap width when
 the heat energy is constant. The results are consistent with the previous theory.

ARTHUR FLEXNER
 The 27th report of the Joint Research Committee of the Institute and Leeds Uni-
 versity. First report on the back-run process for the manufacture of carburized water
 gas. THOMAS HARDIN, et al. *Inst. Gas Eng. Communication* No. 18, 47 pp (1930) —
 This report covers an elaborate test made to det. the efficiency of production of car-
 burized water gas in a modern, automatically operated plant, equipped with generator
 boiler and self-clinkering grate and designed to use the back run principle of heat re-
 generation. Results are given in condensed form: time of test, 60 hrs., coke to gener-
 ator, 151.3 tons, thermal value of coke, 12,920 B. t. u.; oil to carburator, 20,163
 gallons, thermal value, 10,640 B. t. u. per lb., total steam (30 lb. pressure) to generator,
 313,000 lb., av. percentage of steam decompd., 60.3, total carburized water gas made
 (purified) 9,534,000 cu. ft., thermal value of gas, 540 B. t. u., efficiency of production of
 carburized water gas, 69.3%, efficiency of production of blue water gas, (a) accounting
 for extra steam to generator and steam to auxiliary plant, 49.3%, (b) without account-
 ing for extra steam, 55%.

H. L. OLIN
 The 26th report of the Joint Research Committee of the Institution and Leeds Uni-
 versity. Second report on the use of cresote in the manufacture of carburized water
 gas. THOMAS HARDIN, et al. *Inst. Gas Eng. Communication* No. 17, 7 pp (1930) — Gas
 oil, light and heavy cresote, and mixts. of gas oil and cresote were cracked over a temp.
 range of 700-900° in atm. of H₂, H₂ and CO, and H₂, CO and steam. The tabulated
 data indicate the highly unsatisfactory results obtained.

Kind of oil	Therms per gal. of oil	Percent tar	Percent naphthalene
gas oil	1 211	30.1	3.0
light cresote	0 489	67.7	16.6
heavy "	0 167	50.3	18.3
mixt. of 80% gas oil-20% light cresote	1 054	40.1	6.3

H. L. OLIN
 V. D. I. rules for gas-flow measurement with standard nozzles and orifices. K.
 BUNTE and L. ZIFFERER. *Gas u. Wasserfach* 74, 250-4 (1931) — Flow formulas are
 given as well as drawings of standard nozzles and orifices, and practical applications.

R. W. RYAN
 Testing gas-distribution systems with ethyl mercaptan. A. THAU. *Gas u.
 Wasserfach* 74, 247-50 (1931) — A discussion is given of the expts. carried out in 2 Ameri-
 can cities with Et mercaptan to detect leaks in distribution system, and possible appli-
 cation to German conditions.

R. W. RYAN
 Simple flue-gas diagram for excess air. L. ZIFFERER. *Gas u. Wasserfach* 74,
 199-200 (1931) — A simple diagram is given for facilitating the estn. of flue-gas concns.
 for a given gas with various amts. of excess air.

R. W. RYAN

The dew point of the fine gases of solid fuels. W. GRAFTECH *Feuerungstech.* 10, 20-8 (1911) —Gives results that the H content on a wet basis is nearly the same for all solid fuels. This assumption enables him to draw a simple diagram relating dew point of fine gases to the water content of the fuel and to the excess air. FREDERICK W. THURTELL

Coal gas for heating retorts in small gas works. L. SCHUMANN *Gas u. Wasser-fach* 74, 200-71 (1911) —The alternative use of coal gas may permit better balances between gas sent out and coke demand. An economic study is necessary to test the desirability of using coal gas for heating retorts in each case. R. W. RYAN

Use of sewage gas as city gas. W. H. FETTERMAN *Sewage Works J.* 2, 421-31 (1910) —The collection of gas from digestion or Imhoff tanks is not justified in the U. S. as a substitute for city gas because (1) It has a relatively high sp. gr. (0.77) which would necessitate greater expenditure of power for pumping and decrease the capacity of the distribution system, (2) it has a very low H content which gives rise to undesirable flame characteristics, (3) the max. production which could be expected is only 1% of that required for city gas and this would be produced in summer when it is least required, and (4) the enriching quality which sewage gas possesses because of its high B. t. u. value (650 to 800) is obtained at considerably lower cost by the use of bunker oil. F. HERWITZ

Valves for blast-furnace and producer gas. A. JOHANSSON *Jernkontorets Ann.* 114, No. 7, 369-72 (1910) —Detailed description of 600 mm., 400 mm. and 180 mm. valves used in Swedish iron works designed to prevent dust accumulation on the valve seat. I. I. S.

The gas producer. I. Continuous producer. II. Operation and items related to operation. III. Conduction of gas-producer fuel tests. JOHN W. RYAN *Am. Gas J.* 33, No. 2, 47-8, No. 3, 41-5, No. 4, 64, 71-1, No. 5, 45-6, 59 (1910), 334, No. 1, 51-6 (1911). I. I. S.

The disposal of liquor effluents from gasworks. A. C. MONKHOUSE and W. L. THURTELL *Inst. Gas Eng. Communication* No. 21, 12-10 (1910), of C. A. 23, 560-2 —The report is presented under 3 heads: (a) methods of reducing the higher tar acid content of NH₃ liquor, (b) the evaporation of effluent liquor on producer gas and (c) the biological purification of effluent liquors. The principal studies under (a) were directed on the effects of keeping the heavy tars out of contact with the main effluent liquors and thus keeping tar acids out of the latter. Highly promising results with the Cottrell electric frustable tar separator, and with the Cyclone tar extractor are cited. Under (b) the method of evaporating "devil" liquor by spraying it on the grates of producers operating in connection with vertical gas retorts was found to be successful since no fumes or other nuisance was noted. It was estimated that the total "devil" liquor output and 27% of the spent liquor from the NH₃ still could be disposed of in this way. (c) Spent still liquor from the Coventry Gasworks amounting to 7,162,000 gallons by 1910 with an O₂ absorption value 95.0 p. p. m. is being reduced to 9.45 p. p. m. by treatment on bacterial beds. The latter, which consist of graded granite about 5 feet deep, are activated during rest periods with sewage. H. L. OGIN

Catalytic desulfuration and hydrogenation of a primary tar fraction. J. M. PIERCE *Anal. Chem.* 28, 1115-60 (1910) —The S content, existing as H₂S, mercaptans, thiophenes, and thiophene thioles, of a 180-300° fraction of primary tar, was determined, then the tar fraction desulfurized by passing the vapors by means of a current of dry H₂ over oxides previously reduced and heated to 300-350°. Two passages over the catalyst reduced the S content 60%. A second catalyst following the first reduced the base content 61.2% and the tar phenols 13.3%. In hydrogenating this primary tar from which fractions boiling below 180° had been removed previously, by the Bergius process with Fe and Mo oxides as catalysts, there is obtained a benzene, b. p. 60-185°, with a content of phenols and bases reduced 74.0% and 50.3% resp., and wholly sol. in H₂O, which proves the transformation of the 2.8% "sulphur" of primary tar. F. M. SYMMES

The coking industry and its development in relation to the manufacture of iron and steel. JAMES K. DICKIN *J. West Scotland Iron & Steel Inst.* 38, 1st 4, 15-55 (1911). I. I. S.

Influence of water in the by-product coke oven. G. H. FOXWILL *Colliery Eng.* 7, 397-9, 418-10 (1910) —The occurrence of hygroscopic moisture seems to depend on the presence of minute pores in the material. Coke made from coal containing much hygroscopic moisture is very reactive. Water vapor affects the yield of NH₃. Thermal effects are discussed. ANN NICHOLSON HURN

Gas coke as a raw material for household briquets. R. KADA *J. Fuel Soc. Japan* 10, 210-32 (1911); Section 2, 15-7 (in English). —K. made briquets by mixing

sempite with gas coke in various proportions and studied their properties. The equiv. mixt. of semi coke and gas coke produces most suitable briquets which resemble anthracite briquets in their combustibility. K. also suggests the production of coke suitable for household briquets by adjusting the temp. of carbonization. F. I. NAKAMURA

Formation of CS_2 from H_2S and coke (DRAKELEY, BAKER) 18. A new physico-chemical explanation of the formation of humus, peat and coal (ZOLCIN'KI) 8. Beneficial action of brown coal upon the development of cultivated plants (KISSRL) 15. Oven for dry distillation of brown coal (U. S. pat. 1,796,264) 22. Phenols (Brit. pat. 338,638) 10. Apparatus for separating peat fiber from liquids (Brit. pat. 338,547) 23. Continued reduction of ores and hydrogenation of carbonaceous materials (Brit. pat. 339,276) 9. Destructive hydrogenation (Brit. pat. 338,544, 338,576 and 339,317) 22. Purifying bitumen, etc. (Brit. pat. 339,470) 22. Distilling or drying carbonaceous materials (Brit. pat. 338,979) 13

RE UMBERTO, AND VARETON, E. Carburanti e carburazione. Milan: Ulrico Hoepli. 516 pp. L. 50. Reviewed in *Ind. Eng. Chem.* 23, 597(1931)

Method and plant for drying fuels with circulated superheated steam. RUDOLF STERNER and VIKTOR SEITL. Austrian 121,405, Dec. 15, 1929

Low-boiling liquid fuels from coals, tars, oils or bituminous materials. A. ROW-WALTER. Hung. 100,880 Dec. 5, 1929. The raw materials are powdered or vaporized and treated at 350-550° with a mixt. of CO and steam at a partial pressure lower than 1 atm. The low boiling liquid fuels formed are led away in vaporized form.

Motor fuel. I. G. FARRENTINO A-G. Fr. 37,396, Aug. 5, 1929. Addn. to 643,785 (C. A. 23, 1739). In the process of the prior patent the extr. is carried out in steps under conditions which become more and more severe and the residue is afterward submitted to hydrogenation under pressure.

Fuel briquets. GUSTAV KOMAREK (one half to Monarek-Greaves & Co., and one-fourth each to George MacPhail and Charles Coryell). U. S. 1,796,455, March 17. Petroleum coke is used with a coke product from the destructive distn. of coal.

Charcoal for iron metallurgy. E. GARROVITZ. Hung. 101,900, Jan. 7, 1930. Briquets are formed from charcoal dust and lime-contg. water. A g. 22.5 parts CaO is added to 100 parts H_2O , then 80 parts of this solution is mixed with 100 parts of powd. waste charcoal and charcoal dust and the forms are pressed under 2-3 atm.

Distilling fine bituminous material. KOLENVEARDELUNG A-G. Ger. 515,911, Mar. 18, 1924. The finely powd. material is treated with a blast of hot gas so that the distn. begins at above 800°.

Apparatus for continuously distilling carbonaceous materials and cracking or hydrogenating the vapors. "INTERCARBO" SOC. ANON. POUR LA CARBONISATION ET LA TRIÈLEMENT CATALYTIQUE DES COMBUSTIBLES. Ger. 518,403, Mar. 20, 1927. See Brit. 277,404 and 278,041 (C. A. 22, 2655).

Destructive hydrogenation. H. P. DEAN and IMPERIAL CHEMICAL INDUSTRIES, LTD. Brit. 333,640, Oct. 4, 1929. In a described app. for the destructive hydrogenation of liquid carbonaceous material, the H_2 is passed into the reaction chamber through an indirect heat exchanger surrounded by the material and which may be made of or coated with catalytic material.

Destructive hydrogenation. GUTENBERGWERKE OBERHAUSEN A-G. Brit. 339,048, Oct. 8, 1928. Fatty oils such as soy bean oil are converted into products suitable for operation of internal-combustion engines by heating to above their b. ps. in the presence of H_2 at a pressure below 45 atm. (suitably by use of illuminating gas and a catalyst prepd. by pouring dil. Na_2CO_3 over Fe oxide and drying and reducing the mixt.) Use of catalysts contg. Zn and Cu is also mentioned. If desired, the process may be carried out in 2 sep. stages, vapors being formed in the first stage and catalytically hydrogenated in the second stage.

Apparatus for destructive hydrogenation. W. R. TATE, H. P. STEPHENSON, J. F. LERMAN and IMPERIAL CHEMICAL INDUSTRIES, LTD. Brit. 333,479, Nov. 29, 1929. Various details of app. are described, particularly relating to control of the liquid level in a reaction vessel or catch pot.

Apparatus for low-temperature distillation of solid carbonaceous materials such as coal or oil shale. OSWALD HELLER (to Bama Meguin A-G). U. S. 1,796,100, March 10. Various structural details are described of an app. comprising rotating drums and adapted for heating the material undergoing distn. by use of preheated heat-accumulating elements which may have a net like cross section.

Hydrogenating coal, etc. I G PARNFENN A-G Fr 37,334, Sept. 6, 1929. Addn to 666,683 (C A 24, 1201). Other means than that in Fr 666,683 is used to keep the catalyst dispersed in the liquid contained in the reaction chamber, e.g., the catalysts may be used in the colloidal state so that they remain easily in suspension, or the difference in d between the substances to be treated and that of the catalyst is reduced so that the catalyst remains easily suspended. Examples are given.

Preparing peat for briquetting. MADRUCK, GUS FOR MASCHINELLE DRUCKEN-UNGEN IN N. H. Brit 379,970 Aug 23, 1929. Raw peat is mixed with peat dust (preferably in such quantity that the wt of dry material in the dust exceeds that in the raw peat), and the mixt is highly compressed (as by a hydraulic press) to ext water. Numerous details of app and operation are described.

Lignite distillation. H DRACHM. Brit 379,129, Dec 17, 1929. Destructive distn is effected in an app such as described in Brit 270,921 (C A 22, 1073), first at relatively low temp (300-500°) and then at higher temp (suitably 800°), for removal of heavy hydrocarbons and production of a clean and brilliant coke suitable for making briquets (various details of manuf of which also are given). App is described.

Plate drier for lignite. "F. INTRACHT" BRAUNKOHLENWERKE UND BAUKETTFAHRKEN UND MAX MAYER. Ger 515,620, Dec 11, 1929. Addn to 511,711 (C A 25, 2273). Constructional details.

Phenols from industrial liquors. FAITZ HELICH (to Zech Mathias Stinner). U S 1,795,782 March 10. Phenol bearing liquors such as ammoniacal liquors are treated with tar bases such as pyridine or quinoline in order to improve the recovery of phenols.

Preparing highly concentrated ammonia gas from ammoniacal liquor. I. P. TOKIN and I. I. LINER. Russ appl 18,221, May 4, 1927 addn to pat No 3876. Concentrated ammoniacal liquor is dil'd with hot waste water from the distn column. The hot dil mixt is passed through the dissociator for the removal of CO_2 and H_2S in gaseous state. The partly purified water is then freed from NH_3 , H_2S and CO_2 in the distn column.

Purifying benzene. I G PARNFENN A-G Fr 37,363, Aug 23, 1929. Addn to 621,607. The purification of benzene or the intermediate product by a treatment with H under pressure and at a high temp in the presence of catalysts not attacked by S is carried out in a single operation with the hydrogenation of substances such as coal, hydrocarbons or their derivs.

Apparatus for desuperheating steam. SUPPRIMATER CO, LTD, and COMPAGNIE DES SURCHAUFFEURS. Brit. 379,575, March 19, 1930. Structural features.

Gas production. A. S. RAMAGE. Brit. 378,005, Aug 21, 1929. Olefins, terpenes or acetylenic hydrocarbons, substantially free from paraffinic hydrocarbons, are passed, at a temp of about 420-700°, over iron, with 20-30% or more of steam, producing mainly H_2 , CO and C_2H_4 , and leaving the iron as metal. Various details of app and procedure are described. Cf. C A 24, 2866.

Fuel gas. HERMANN HENDEL. Austrian 121,091, May 15, 1930. In the manuf. of fuel gas by degasifying coal and converting the residual coke into water gas, the degasification is effected by leading a part of the water gas through the degasifying retort and another part around the retort, the latter part being afterward reheated and returned to the water-gas generator with superheated steam. The calorific value of the gas may be regulated by controlling the amt of water gas admitted to the degasifying retort, complete degasification being in any case ensured by regulating the temp of the water gas so admitted, and also the temp and the amt of water gas led around the retort. Cf. C A 24, 228.

Fuel gas. ANTON KRAATKY. Austrian 121,099, Aug 15, 1930. Gas for heating or for driving internal combustion engines is prepd by the action of steam at 500-1000° on atomized or gaseous or vaporized hydrocarbons in a chamber heated electrically to approx the temp of the steam. A/c vapor may be mixed with the steam. App is described.

Fuel gas from bituminous materials. ARTHUR H LYNN. Ger 515,982, Feb 22, 1927. Details of retorts and method of distg are given.

Coal gas. MARIA T. STRACHE. Austrian 121,510 and 121,512, Oct 15, 1930. Addns to 118,062 (C A 24, 4021). In the method of Austrian 118,062, uniform distribution of the fresh fuel and tar is effected by feeding the fuel to the generator through the tar, which may act as a seal for the generator. App is described (121,510). Addnl tar may be added with the fresh fuel, besides the tar produced in the gasification. Low-temp tar, generator tar, wood tar, or lignite tar may be so added. Mineral oils or their distn products or resins may be added also (121,512).

Desulfurizing coal-distillation gases, etc. ISRYAN HUNYADY and KARL KOLLER. Ger 518,431 May 11, 1928. The gases are freed from tar, NH_3 and CN compds, and

then treated with an aq. suspension of $Mn(OH)_2$. The sludge of MnS and S so obtained is withdrawn and treated with SO_2 to produce a suspension of S in an aq. soln of $MnSO_4$. The soln is sepd. and treated in turn or simultaneously with the gases to be purified and with SO_2 whereby $MnSO_4$ is converted into $Mn_2S_3O_4$, which is stable to H_2S . The soln of $MnSO_4$ is then boiled, yielding $MnSO_4 \cdot 5H_2O$ and SO_2 . The $MnSO_4$ is treated with NH_3 and O to regenerate $Mn(OH)_2$.

Oil gas. FRANK J. NOLAN. U. S. 1,796,299, March 17. A mixt. of vaporized hydrocarbon material and steam is introduced to a zone contg. a plurality of individually spaced metallic catalysts such as an alloy of $Al 50$, $Ni 15$, $Cu 10$, $Ti 10$, $Ag 2$, $Mg 10$ and $Bi 3\%$, each spanning the zone and heated to above 540° and the mixt. is passed over the catalysts without further addn. of steam and the gas produced is withdrawn for collection. App. is described.

Oil gas. A. S. RAMAGR. Brit. 338,904, Aug. 21, 1929. Materials such as heavy mineral oils are passed with 1-2% of steam, over ferric oxide at a temp. of about $420-540^\circ$ and the (largely olefinic) gases thus formed are passed, with 10-40% of steam, over finely divided iron at a higher temp. (suitably about $600-800^\circ$), producing a gas contg. H_2 , CO , C_2H_4 , C_2H_2 and some aromatic hydrocarbons. Motor fuel is sepd. from the gas and various details of app. and procedure are described.

Water gas. A. F. KUNBERGER (to Humphreys & Glasgow, Ltd.). Brit. 338,804, March 2, 1929. H_2 produced by passing steam over hot iron is superheated and carbureted and then passed with excess steam through the upper part of the fuel bed (the whole of which is used simultaneously in generating water gas by "up-running") and the mixed gas is taken off above the fuel bed. Various details of app. and operation are given.

Water-gas producer. MOTOR FUEL PROPRIETARY, LTD. and W. JOHNSON. Brit. 338,877, Aug. 26, 1929. Various details are described of an app. having an inner generator located within an annular generator, charges in each generator being independently gasified.

Water-gas producer for the continuous distillation of powdered fuel. JULIUS PRITSCH A. G. Fr. 37,192, July 4, 1929. Addn. to 658,595 (C. A. 23, 5377).

Apparatus for making water gas from pulverulent fuel. HARALD NIELSEN and BRYAN LAROC. Ger. 518,427, May 31, 1927. See Brit. 299,485 (C. A. 23, 3333).

Gas producer. S. I. R. I., Soc. ITALIANA RICERCHE INDUSTRIALI. Brit. 338,911, Aug. 27, 1929. See Fr. 680,849 (C. A. 24, 3932).

Gas producer (suitable for generating gas from solid fuels on automobiles). WM. W. ODELL and EARL N. PERCY. U. S. 1,795,670, March 10.

Combined gas-producer furnace suitable for heating metals. R. A. HADFIELD and R. J. SARGANT. Brit. 338,893, June 26, 1929.

Gas-generating plant. ADOLF WIEGEL. Ger. 515,083, May 10, 1925. Details of a surface cooler for the gases produced are given.

Apparatus for producing fixed combustible gas from fluid fuels such as oil. ORRELL O. THIRING (to General Oil Gas Corp.). U. S. 1,796,733, March 17. Numerous structural details are described.

Enriching combustible gases. HERMAN A. BRASSETT and CHARLES W. ANDREWS (to H. A. Brassett & Co.). U. S. 1,795,829, March 10. A gas to be enriched, together with a regulated proportion of air, is passed through a mass of partly consumed coke at high temp., then through a mass of coke at still higher temp. which will effect formation of CO from CO_2 , and a controlled proportion of preheated air is introduced laterally into the last mentioned coke mass, and the gas is then passed through a mass of coal at a lower temp. to effect enrichment. App. is described.

Washing flue gases, etc. P. J. ROBINSON. Brit. 338,492, May 18, 1929. Gases such as flue gases are passed through a centrifugal dust separator, adjusted to leave a regulable small proportion of fine dust in the gases, and are then subjected to steam or fine water sprays to wet the remaining fine dust particles and effect oxidation of SO_2 to SO_3 , after which the gases are caused to impinge upon baffle plates and are washed with a heavy shower of water to remove the dust and acid. Various details of the app. (which may be formed of Ni or Cr steel resistant to corrosion) and of operation are described.

Regenerative heat-storing device, especially for preheating fuel gas. AKTIEBOLAGET LJUNGSTROMS ANSTADEN. Ger. 515,870, Nov. 15, 1927.

Refining tars, oils, etc. I. G. FARRENING A.-G. (Carl Krauch and Mathias Pier, inventors). Ger. 515,523, Aug. 15, 1925. The tars, oils, distn. products, residues, etc., are treated with H or H -evolving materials at high temps. and pressures, optionally, in the presence of catalysts. The material to be refined by this process is intimately mixed with up to 30% peat or lignite or a mixt. of both. In an example, lignite gener-

ator tar is mixed with 20% crude lignite and heated in a Mn bronze-lined high pressure reaction chamber to 450°, and subjected to the action of H₂ at 200 atm. A 20-35% yield of benzene results. Further examples are given. Cf C. A. 24, 4378.

Distillation of tar. ALEXANDER A. MACCUBBIN and JOSEPH ZAVERTNIK (to The Barrett Co.) Can 309,218, Mar 10, 1931. A method and app are specified in which tar is distd at atm pressure to produce distillate oils and a pitch residue. The pitch residue is withdrawn during distn at approx the temp of the first distn and immediately subjected to a high vacuum to effect further distn of high boiling oils by the self contained heat of the pitch. Cf C. A. 24, 2585.

Removing acid constituents from low-temperature tar or its fractions. CURT BUNGE and FORSCHUNGSINSTITUT FÜR BERGWERKE- UND SPRENGSTOFFCHEMIE SOWIE VERWANDTE GEBIETE. Ger 518,210 Sept 13, 1925. An aq emulsion of the tar or tar fraction is treated with alkali soln. The emulsion may be prepd with the aid of soap, and the treatment with alkali soln may be effected in a no. of stages.

Chamber oven for the production of gas and coke. C. OTTO & Co G. M. & H. Ger 515,978, Oct. 20, 1926. Details of heating the distn gas to produce gas of better illuminating quality are given. Cf C. A. 24, 4379-5139.

Coke oven. RUDOLF WILHELM (to Arnold Beckers). U. S. 1,795,324, March 10,

Coke oven. KOKSOFFENBAU UND GASVERWERTUNG A.-G. Ger 512,334, Nov 22, 1927. The heating channels are lined with steel or other heat resistant metal.

Regenerative coke ovens. COLLIN & Co. Fr 37,387, Sept 2, 1929. Addn to 576,379.

Regenerative compound coke oven with twin outlets. C. OTTO & Co G. M. & H. Ger 515,849, Aug 18, 1929. Addn to 503,895 (C. A. 24, 5957).

Coke-oven battery with horizontal chambers. THE KOPFERS CO. Ger 518,312, Dec. 15, 1928. See Brit. 302,365 (C. A. 23, 4331).

Heating wall for coke-oven chambers. KOKSOFFENBAU UND GASVERWERTUNG A.-G. Ger 518,355, April 29, 1926.

Coke oven (door construction, etc.). R. & J. DEMPSTER, LTD., and C. E. HOLT. Brit. 339,557, Feb 19, 1930.

Coke-oven door. HEINRICH SCHWARZ. U. S. 1,795,515, March 10,

Coke-oven door. E. WOLFF. Brit. 338,459, Oct. 9, 1929.

Coke-oven cover. ARTHUR KILLING and WILHELM ELBERT. Ger 515,528, Aug 19, 1928.

Method and plant for drying coke. WESTFÄLISCHE MASCHINENBAU-GESELLSCHAFT. Ger 518,313, June 17, 1930.

22—PETROLEUM, LUBRICANTS, ASPHALT AND WOOD PRODUCTS

W. F. FARACHER

Petroleum mine in the old Gromni District. L. LUTOVINOV. *Azerbaidzhanskoe Neftyanoe Khozyaistvo* 1931, No 1, 57-63.—A complete description is given of the proposed method of mining oil similar to that adopted in Pechelbren (France).

V. KALICHEVSKY

Determination of the toluene content of a Mid-Continent petroleum. JOHANNES H. BRUN, R. T. LESLIE and SYLVESTER T. SCHICKTANZ. *Bur Standards J. Research* 6, 363-7 (1931).—Com petroleum fractions from an Okla. crude oil were fractionated into 1° cuts and tested for toluene by nitrating to 2,4-dinitrotoluene by means of a nitrating method which was found not to attack the other constituents of the fractions. Toluene was found in all the cuts boiling between 99° and 112°. The max concn (32.2%) was found in the fraction boiling between 107° and 108°. Total toluene present amounted to 1/4 of 1% based on the crude petroleum. D. F. BROWN.

Separation of normal octane from petroleum by distillation and crystallization. R. T. LESLIE and SYLVESTER T. SCHICKTANZ. *Bur Standards J. Research* 6, 377-86 (1931).—A fraction b 100-130° from an Okla. crude oil was subjected to an interlocking process of distn. and crystn. to sep octane. The method and app are described in detail. The b. p., *n*_D, *n*_F, refractive index and mol. wt. of the product are given and compared with the properties of a carefully purified sample of synthetic octane. It is estd that not less than 1% of octane was present in the oil investigated. D. F. BROWN.

The acids of Baku petroleum. A. E. CHECHIBADIN. *Compt. rend. acad. sci. U. R.*

S S 1930A, 332-4. A summary is given of work done on petroleum and naphthenic acids. The sepn of the acids from the lower fractions as amides showed that fatty acids with 7 C atoms were present. Bromination and sepn of HCOOH from the hydroxy-acids proved from the bromoacetic acids indicated that the carboxyl group is bound to primary radicals as well as to secondary and tertiary radicals. By sepg the acids into cyclic and fatty acids by means of their Cd salts, it was found that the petroleum acids b 21-24 consist principally of fatty acids. The content of cyclic acids increases regularly with the b p, so that acids b above 268° (C_{14} acids) belong almost exclusively to the cyclic series. The consts of the so-called lower naphthenic acids do not agree with those of acids with 6- and 6-C rings. The C_7 acid, heptanaphthenic acid, is really a mixt of fatty and naphthenic acids. L. JACOVITZ

The refining of light petroleum distillates. I W SUREL. *Chem. Eng. Mining Rev.* 23, 171-8 (1929). The predominating impurities in straight run distillates are S compounds which in unduly amounts cause corrosion, particularly of the brass and Cu fittings of the carburetor. Data are given to show that the S content of distd fractions from any oil increases with rise in b p. The chief chem refining processes are described briefly: 1. the NaOH , Na plumbate, H_2SO_4 , Frasch and Na hypochlorite processes. In the refining of cracked distillates desulfurization with the simultaneous removal of gum forming diolefins is the objective. Acid treatment combined with sweetening is usually given these cracked distillates. A batch and a continuous treating system are described. D. P. BROWN

Report of Subcommittee XXVIII on autogenous ignition of petroleum products. A H NUCKOLLS et al. *Proc. Am. Soc. Testing Materials* 30, Pt. 1, 78-112 (1930). As a result of cooperative tests on aviation gasoline, ordinary gasoline, Standard solvent kerosene, motor oil, I O , acetone, CS_2 and heptane, the results of which are given, minor changes in the tentative standard D 294-28T are suggested. It is recommended that the test be adopted as standard with these revisions and that further cooperative tests be made. D. L. BROWN

Atmospheric and vacuum distillation of Mid-Continent crude oil. JOHN PARKROSS. *Refiner and Natural Gasoline Mfr.* 10, No. 3, 69 (1931). The processes and equipment applicable to atm and vacuum distn are discussed. J. L. FARRER

Physical properties of mineral oils and their investigation. I. TIHAMAR ASZAGYAR. *Művelőd., Epítészművelőd. és Könyv. Hír-Füzetek* 8, 27-38 (1931). A general description. S. S. DE FINÁLY

Refining shale oils with silica gel and with bauxite. II. Filtrations in liquid phase. B. SALADINI. *Industria Chimica* 5, 1452-7 (1930) cf. *C. A.* 24, 2270, 25, 2276. Shale oils contg as high as 9.5% S were filtered through silica gel, bauxite, mixts of these, and also mixts of each with decolorizing materials. Silica gel alone or with decolorants gives the best results removing up to 30% of the total S, whereas bauxite removes up to 18-20%. Mixts contg bauxite, as well as the material itself, cause trouble in filtering, so that silica gel is preferable. A. W. CONTIERI

Black shale deposition in central New York. EDWARD W. HARR. *Bull. Am. Assoc. Petroleum Geol.* 15, 165-91 (1931). Lab distn tests of N. Y. black shales show the amt of petroleum increases to the west, but with no corresponding increase in content of volatile matter, indicating that the petroleum was derived from org. constituents. The bituminous content of these shales seems to be directly related to the type of decay compound rather than to the type of org. material, and this particular type of decay existed only where the water was truly saline and toxic conditions were present. ALDEN H. EMERY

Discovery of oil in White Point gas field, San Patricio County, Texas, and history of the field. W. ARMSTRONG PRICE. *Bull. Am. Assoc. Petroleum Geol.* 15, 205-10 (1931). ALDEN H. EMERY

Heavy Dutcher oil in Bristow district, Oklahoma. CHARLES G. CARLSON. *Bull. Am. Assoc. Petroleum Geol.* 15, 211-3 (1931). The gravity of most of the Dutcher oil produced in the Bristow district is 29-34° Be. A new well gives oil of 18.3° Be, asphalt 92 S 0908, gasoline 6 and kerosene 2%. ALDEN H. EMERY

Development and production history on the Salt Flat and other fault fields of east central Texas. H. B. HILL, E. V. H. BALSERMAN and C. B. CARPENTER. *Bur. Mines, Rept. of Investigations* 3059, 46 pp (1931). ALDEN H. EMERY

New productive horizon in California. WALTER STALDER. *Bull. Am. Assoc. Petroleum Geol.* 15, 201 (1931). ALDEN H. EMERY

Can absence of edge-water encroachment in certain oil fields be ascribed to capillarity? J. VERSLUY. *Bull. Am. Assoc. Petroleum Geol.* 15, 189-200 (1931). Capillarity cannot hold the oil outside the depleted area around the well and cannot prevent

the edge water from moving toward the well. In the depletion zone, when the formation is an unconsolidated sand, 2 conditions are possible: the funicular, in which oil and generally gas flow freely, and the pendular, in which gas flows freely and may be followed by oil. If the oil-bearing formation is consolidated and the interstices are pores with narrowings, the Jamin effect would not stop the flow altogether because the gas would diffuse through the films of oil which shut off the pores in the narrowings. Light oils would evaporate from one film and condense on the next nearer to the well.

ALDEN H. EMERY

Some effects of metamorphism on certain debris in source rocks. TAISSIA STADNICHENKO *Bull. Am. Assoc. Petroleum Geol.* 15, 161-4(1931)—Geodynamic alteration (carbonization) of the plant substances in the source rocks of petroleum results in raising the temp. of fusion and volatilization and in change in optical characters from isotropic to anisotropic.

ALDEN H. EMERY

Concrete coverings for pipe lines. J. F. HOUGH *Oil and Gas J.* 29, No. 35 62, 140, 141(1931)—From tests carried out under the most severe conditions it was found that concrete-covered pipe resisted corrosion satisfactorily. The best results were obtained by using a concrete with less than 5 gal. of water per sack of cement. More water causes the concrete to disintegrate. Pressures of 1000 lb. per sq. in. did not give pressure cracks in the concrete coverings.

J. R. STRONG

The compression of hydrocarbon gases. I. N. BEALL *Refiner and Natural Gasoline Mfr.* 10, No. 2, 99, No. 3 154(1930)—The mathematical development of formulas pertinent to compression of gases is described. Calculation of the percentage of liquefaction on compression, molal vol. of hydrocarbons, importance of compression before condensation for certain types of oil absorption plants, compression at high pressures and multiple stage compression are discussed. All gases become more compressible as their temp. is lowered and deviations from Charles' and Boyle's laws reach a maximum at equilibrium.

J. L. ESSEX

Safe-guarding cracking operations from accident. E. M. MATSON *Refiner and Natural Gasoline Mfr.* 10, No. 1, 97, No. 2, 122, No. 3, 164(1931) cf. *C. A.* 25, 1066—The hazards of water in a cracking system and the dangers arising from the corrosion of the equipment are discussed. Methods of preventing both occurrences are considered.

J. L. ESSEX

Manufacture of cracked tractor fuel. V. VALGIS *Azerbaidzhan'skoe Neftyanoe Khozaystvo* 1931, No. 1, 64-73, cf. *C. A.* 24, 2588—Exptl. work is described.

V. KALICHEVSKY

Use of the new natural-gasoline specifications. S. S. SMITH *Refiner and Natural Gasoline Mfr.* 10, No. 2, 90(1931).

J. L. ESSEX

Progress in distillation and fractionation. EARL PETTY *Refiner and Natural Gasoline Mfr.* 10, No. 2, 67(1931)—The changes in distn. procedure due to the use of stabilizers are discussed and the use of vacuum in the distn. of heavier lubricating-oil fractions is described.

J. L. ESSEX

Pressure-distillate stabilization and gas-recovery systems. A. W. BURKET *Refiner and Natural Gasoline Mfr.* 10, No. 3, 75(1931)—The essential factors are given by means of which the basic app. can be modified to suit any particular plant. For every vol. of propane removed from pressure distillate, 3.5 or more vols. of butane can be introduced without changing the vapor pressure. A complete description of the equipment necessary for combined stabilization and recovery is given.

J. L. E.

Chemistry of fuels: greatest problem before refiner and motor builder. EARL BARTHOLOMEW *Natl. Petroleum News* 23, No. 11, 72, 76, 79, 81, 82, 84, 85(1931)—The 2 variables most affecting the relative knock ratings are temp. and air-fuel ratio. Tests have shown that the single-cylinder test conditions give too high an antiknock value for benzene and some cracked gasoline.

R. W. KELLY

Chemistry of the doctor sweetening process. ARTHUR LACHMAN *Ind. Eng. Chem.* 23, 354-7(1931)—L. verifies the 2 views that the PbS formed during the sweetening reaction reacts to form more plumbite and also acts as a catalyst when a sour gasoline is blown with air in the presence of Na plumbite soln. L. claims that Na₂S is also formed, which then oxidizes to Na₂S₂O₃. The sensitiveness of the doctor test was found to be 0.00006 molar concn. of mercaptan. The molar concn. of mercaptan in an av. sour gasoline was 0.01. Mercaptans can be oxidized with air alone in the presence of NaOH soln., or in the presence of fine inert solids, provided that the air stream is broken up into small bubbles. The quantity of Na₂S in alk. soln. contg. suspended PbS should be kept at a min. Air blowing assists in fulfilling this requirement. Finely divided free S should be added in the quantity detd. by analysis of the distillate for mercaptans.

J. L. ESSEX

decanted and centrifuged, 0.020%; with 10% earth and decanted, 0.018%; and with 10% of earth, decanted and centrifuged, 0.018% R. SANSONE

Raw materials and production of mineral lubricating oils. M. FREUND *Magyar Mérnök Építészegylet Közlönye Havi Füzetes* 8, 1-10(1931)—A general description of present production methods S. S. DE FINÁLY

Principles of chemical investigation of mineral lubricating oils. M. FREUND *Magyar Mérnök Építészegylet Közlönye Havi Füzetes* 8, 18-25(1931)—A general description of research methods S. S. DE FINÁLY

Regeneration of used lubricating oils. F. SCHNITZER *Azerbaidzhanskoe Neft yano Kheymalstvo* 1931, No. 1, 107-8— Na_2SiO_3 , NaOH and bleaching earth are almost equally satisfactory V. KALICHINSKY

Vacuum. GERRITT VAN DEN BERG *Refiner and Natural Gasoline Mfr* 10, No. 2, 88(1931)—A few schemes for preventing leaks are outlined J. L. ESSBX

Report on coöperative work on the separation of cut-back asphalt. W. H. FULWEILER *et al Proc Am Soc Testing Materials* 30, Pt 1, 801-4(1930)—Coöperative results from the use of 2 different methods for sepg cut back asphalts to det. the amt and character of both the base and the solvent are given. The first method is described in tentative specification D 20 28T. The second starts with atm distn and completes the distn in a vacuum of not more than 10 mm of Hg. The second method appears to have certain disadvantages. The work is to be continued and certain modifications suggested are to be tried D. F. BROWN

Utilization of wood waste in the chemical industries. B. B. FOGLER *Am Soc Mech Eng*, Advance paper, Oct 16-17, 1930, 4 pp.—Principal uses, other than for fuel for wastes from saw-mill and wood working industries are found in the following chem industries: destructive distn, manu of pulp, paper and building boards and those industries consuming large quantities of wood flour, principally the linoleum, plastics and explosives industries. The paper is intended to present a comparative summary of several outlets which chem industries afford for salvage of waste from wood working industries. E. I. S.

Wood chips. I. Motor oils from wood. OSKARI ROUTALA *Acta Chem Fennica* 3, 115-8(1930)—A review is given of the methods for cracking oils to obtain the lower fractions for use in gasoline motors. Lignite at 150-200 atm and 450-80° yields about 60% of oily products, which yield considerable gasoline on cracking. When sawdust is treated with Cu, Fe and Co salts, and then with NiH_2O_2 or NaOH, an impregnation with metallic hydroxides occurs. The latter are reduced to the free metal on treatment of the sawdust with H under pressure at 400-450°. Ni formate, NiH_4 molybdates and chromates are also used as catalysts, a complete conversion of the sawdust to oily products resulting. At 200°, 100% conversion is obtained. On distn up to 350° the product gives water 4.45, oil 35.53, residue 52.10, gas 6.73 and loss 1.17%. Of the oil fraction, 74% distd below 300°. Sawdust at 450° and 100 atm with Zn and Cu catalysts gave 33% conversion to oils, of which 40% distd below 200°. Forty five % of the weight of the wood was water-sol., with low acid but high MeOH content S. A. KARJALA

Dust explosions, with special reference to wood-working industries (BROWN) 24. The radioactivity of oil-well waters (VERNADSKII) 3. Cracking oil by electricity (ROWLAND) 4. The content of Ra in oil-well waters in the district of Grozny (KILOPIN, NIKITIN) 3. Washing and bleaching clays of Azerbaidzhan (KOVALEVSKII) 18. Apparatus for low-temperature distillation of oil shale (U. S. pat. 1,790,100) 21. Pump and high-speed emulsifier system for bituminous materials (Brit pat 338,496) 1. Refining oils (Ger pat 515,523) 21. Detergents and lubricants containing fatty acid salts (Brit pat 338,910) 18. Gel-like benzene (Hung pat 101,509) 13. Gelatinizing organic liquids such as benzene (Brit pat 339,360) 13.

Converting petroleum oils. ROY CROSS (to Gasoline Products Co.) U. S. 1,790,507, March 17. Oil is raised to a cracking temp without substantial conversion (suitably by heating in a pipe coil in a furnace), the oil is collected in a substantial body and lighter fractions are distd off in a vaporizing stage with refluxing and condensation of the vaporized material. Unvaporized products are directed to a stagnant pool adapted to be heated to facilitate settling of free carbonaceous matter, the heavy settlings are withdrawn from the bottom of the pool and the liquid, relatively carbon-free oil, is withdrawn from a vertical restricted column in the pool so that turbulency in the oil body is prevented. App is described.

Distillation of petroleum oils. ALEXANDER A. MACCLUBBY and JOSEPH ZAVERTNIK (to The Harrett Co.) Can 309,219, Mar 10, 1931. A method and app are specified for distn of petroleum at atm pressure to produce distillate oils and residual oils. The residual oil is withdrawn during distn while it is at approx the temp of the first distn and immediately subjected to a high vacuum to effect further distn of high boiling oils by self contained heat of the residual oil. The residue from such distn is passed to a coke still maintained under high vacuum and distd to coke, and the distillate is condensed under a high vacuum.

Apparatus for chilling liquids such as petroleum distillates to separate solids from them. HENRY TORRANCE (to Carbondale Machine Co.) U S 1,700,772, March 17. Numerous structural details are described.

Dehydrating emulsions such as those of petroleum. HAROLD C. EDDY (to Petroleum Rectifying Co. of Calif.) U S 1,706,750, March 17. Dispersed particles of the emulsion are mechanically agglomerated (suitably by the action of extensor or steel wool) and the emulsion is subjected to the action of an elec field. App is described. Cf C A 25, 409.

Hydrocarbons. GENERAL TECHNICAL CO., LTD. Fr 695,185, Aug 20, 1929. Bituminous and asphaltic compds, petroleum exts, tar exts and natural waxes are submitted to a temp of 370-440° under ordinary pressure in a retort provided with a system permitting the vapors to pass out to the condensers at a temp between 170° and 270° the speed of distn being preferably such that the hourly amt. of distillate is always above 10% of the wt. of the capacity of the retort. Lighter hydrocarbons such as gasoline and kerosene are obtained.

Solidifying hydrocarbons. ROGER HOLTMANN. Fr 695,194, Aug 22, 1929. Hydrocarbons, such as gasoline, petroleum and mazout, are solidified by incorporating therein before or after heating, a binding agent, such as resin, suitably powd. and then adding while mixing milk of lime. The product is molded under pressure and hardened by immersion in water which preferably contains gallic acid.

Cracking hydrocarbons. I. G. FARABYND A-G. Fr 37,101, July 4, 1929. Addn to 653,906 (C A 23, 5312). Hydrocarbons of high b p, such as mineral oils, tars, etc. are cracked under such conditions that a limited formation of benzine results, and the residue is afterward submitted to a hydrogenation under pressure, the untransformed fraction being submitted afresh to a cracking. Cf C A 25, 1666.

Cracking hydrocarbons. I. G. FARABYND A-G. Brit 339,274, May 30, 1929. Cracking of hydrocarbons in the gaseous or vapor phase is effected in the presence of a catalyst of pieces of Si (which may be molded from powd. Si and may be placed in a quartz tube) or Si deposited on a carrier. The material may be mixed with N, CH₄, water vapor, etc., and examples are given of the cracking of a mixt. of C₁₂H₂₂ and C₁₄H₃₀ at 700° to form C₂H₄ and C₃H₈, and of the cracking of pentane, ligroin, paraffin and petroleum oils. It is stated that no carbon is deposited on the catalyst.

Cracking oils. CHARLES B. BUEGER (to Gulf Refining Co.). U. S. 1,706,138, March 10. Vapors are withdrawn under pressure from a confined body of oil heated under pressure, the pressure on the vapors is released and they are condensed, and the confined body of oil is replenished by oil heated under pressure and introduced under the surface of the body of oil while simultaneously permitting vapors from the preheated replenishing oil to mingle directly with vapors above the confined body of oil. Liquid oil is withdrawn from the main body of oil and mingled with the replenishing oil fed to the preheater. App is described.

Cracking oils. PANHANDLE REFINING CO. Brit 339,291, June 29, 1929. Oil is preheated in a pipe coil and is then passed into a vaporizing zone, whence vapors are passed to a cracking coil, and the cracked vapors are fractionally condensed in a series of dephlegmators and fractionating columns by indirect heat interchange with make-up oil or condensate. Various details of app and operation are described. Cf C A 24, 3637.

Cracking oils. ROBERT E. WILSON and RICHARD J. DEARBORN. Can 309,377, Mar 17, 1931. A process and app for converting higher boiling oils into lower boiling oils are specified. A stream of oil is heated in a coil under high pressure to a temp sufficient to effect conversion, the heated oil is discharged into a conversion chamber maintained under a reduced pressure but substantially above atmospheric, the vapors thus formed are sep'd into several fractions by condensation, the residues from the conversion chamber are expanded into an expansion zone under a pressure below that in the chamber, a portion of the condensates resulting from the first-mentioned sep'n is returned to the heating coil at an intermediate point, another portion of the condensates is delivered into said expansion zone; fresh charge is introduced into said expan-

sion zone, and a fraction of the oil collected in said expansion zone is passed to the heating coil Cf C A 24, 2591

Vaporizing device (suitable for use with automobile engines) for catalytic decomposition or cracking of liquid fuels. D BALACHOWSKY, P CAIRN and M LIPP Brit. 339,251, July 26, 1929 Various structural details are described

Refining hydrocarbon oils. ALEXANDER S RAMAGE (to Gyro Process Co) U S 1,796,621, March 17 S and Cl are removed from alc sol hydrocarbon materials such as olefin and naphthene hydrocarbons or crude benzene by dissolving the materials in an alc soln of an alkali metal hydroxide, permitting the soln to stratify and separately removing the purified oil and the alc alk soln contg impurities

Purifying hydrocarbon oils with liquid sulfur dioxide. PAN AMERICAN PETROLEUM Co Brit 338,452, May 7, 1929 Materials such as lubricating oil kerosene or various cracked distillates are treated at a temp below 0° with liquid SO_2 and, after sepn the partly purified oil contg SO_2 is further treated, also at a temp below 0° , with fuming H_2SO_4 and the purified oil is sepd from SO_2 and sulfo acids Numerous details of app and procedure are described Brit 338,453 relates to a generally similar process in which however, the H_2SO_4 used is not specified as being fuming acid Brit 338,484 describes a process of treating crude petroleum shale or cracked oils with lumping or coned H_2SO_4 in the presence of liquid SO_2 (simultaneously or previously added to the oil) preferably at a temp not greater than the b p of SO_2 thus freeing the oils from polymerizable, unstable and S bearing hydrocarbons Various details of app and procedure are given

Treating hydrocarbon oils with aluminum chloride. JOHN L COOLFEY (to Standard Oil Co of Calif) U S 1,795,761, March 10 A hydrocarbon oil is mixed with AlCl_3 in a flowing stream at a temp sufficient to cause reaction and formation of a relatively low-boiling product, the stream is passed to a flash zone in which vapors of low b p are sepd from residual oil and AlCl_3 , and a stream of the vapors and a sep stream of the oil and tar are passed to a conversion zone in which the still heated vapors are again brought into contact with the oil tar mixt for agitating it and effecting further decompn App is described

Apparatus for continuous vacuum distillation of hydrocarbon oils. KLAUS DRIEGER 515,704, Nov. 1, 1927 Details of arrangement

Separation of mixtures of hydrogen and hydrocarbon gases. W K LEWIS (to Standard Oil Development Co) Brit. 338,719, Nov 12, 1928 See Can 305,731 (C A 25, 783).

Purifying mineral oils. ALLGENEINE GES FÜR CHEMISCHE INDUSTRIE Brit 339,579, July 22, 1929 Material such as the total distillate from crude Persian petroleum up to 400° is treated with liquid SO_2 without previous fractionation, or the distillate may be freed from the benzine fraction b up to 100° before treatment The raffinate and ext are freed from SO_2 and fractionated Cf C A 24, 3109

Adsorbent clay for decolorizing oils. ROBERT E MANLEY and MERTON L LANGWORTHY (to Texas Co) U S 1,796,709, March 17 The clay is treated with a H_2SO_4 soln of about 10% strength to improve its decolorizing properties

Oven for dry distillation of bituminous shales, brown coal and similar materials. JOHAN G GRÖNDAL and CARL L CARLSON (to Patentaktiebolaget Gröndal-Ramén). U. S. 1,796,264, March 10 Structural features of an app with a perforated traveling grate

Destructive hydrogenation. J FOHLEN, Brit. 338,544, June 21, 1929. In the production of light hydrocarbons from materials such as anthracites, other coals, peat wood, petroleum, asphalts or various org vegetable or animal materials (or from products of pyrogenation of these), nascent H is produced from reacting material introduced independently into the reaction vessel (an example being given of the treatment of shale oil mixed with Fe and Cu shavings, introduced into an autoclave at 300° and 300 atm pressure contg an aq soln of a halide such as AlCl_3 , 3NaCl for production of light oils) A reaction of water or steam on coke may also be used for obtaining nascent H Cf C A 24, 1212

Destructive hydrogenation. F UDDER Brit. 339,317, Oct 3, 1928 Low-boiling hydrocarbons are produced from carbonaceous materials such as oils, bitumens, carbon, wood, etc., by heating them with finely divided or spongy iron and water in an autoclave to 400 – 500° with development of 200 atm pressure during the reaction from the vapor pressure of the water, etc., present (the water reacting on the Fe to form H) The gases generated can be used to regenerate the Fe from the Fe oxide formed in the reaction and hydrogenating catalysts such as Ni, Cu, ZnO or Zn chromate or splitting

Co of Ind) U S 1,790,857, March 17. A fluid lubricating oil is used with an addn. (suitably in the proportion of about 7%) of a relatively volatile and substantially non-inflammable diluent such as CCl_4 .

Lubricant suitable for use in drawing and polishing metals. HERBERT L. JOHNSON (to Standard Oil Development Co) U S 1,795,491, March 10. A salt (such as the Na salt) of a substantially water sol acid obtained from the sludge formed by treating petroleum lubricating oil with H_2SO_4 is heated together with water and a fatty oleaginous material such as tallow under such conditions as to effect distribution of the fatty material in the water.

Lubricating oil. WM O STEVENS (to Anti Hot Box Co) U S 1,796,310, March 17. Dried wood ashes are immersed in lubricating oil in order to improve the oil, and the ashes and oil are then sepd.

Lubricating oils from petroleum. RED RIVER REFINING Co, Inc. Ger 515,840, Mar 17, 1923. App for distg the petroleum *in vacuo* at 1-5 mm pressure is described.

Apparatus for filtering lubricating oil. WM W NUGENT U S 1,796,613, March 17. Structural features.

Bag filter suitable for filtering oil. WM W NUGENT U S 1,796,632, March 17. Regeneration of used lubricating oils. I THAMM HUNG 100,735, Dec 20, 1929.

An aq emulsion is made of the changed part of the oils which can be sepd by sedimentation from the unchanged lubricating oil. Emulsification may be accelerated by adding carbonates or oxides of earth metals to the oil. The treatment with steam emulsifies the changed oil particles and carries away the fuel particles which were mixed with the oil. Grease can be made of the sepd emulsion and the regenerated oil may be filtered and used again for lubricating.

Wood distillation plant. JOHN T MYERS U S 1,795,404, March 10. Various details of construction are described.

23—CELLULOSE AND PAPER

CARLETON E. CURRAN

Some new processes for the separation of fibrous cellulose from plant substances. J. MELROSS ARNOT. *World's Paper Trade Rev.* 95, 961-4 (1931).—An address in which are briefly discussed some of the most recently proposed pulping processes. A. P. C.

Absorption of water vapor by cotton cellulose. ROBERT H. PICKARD. *J. Am. Chem. Soc.* 53, 1610-1 (1931).—Polemical with Pidgeon and Maass (*C. A.* 24, 1974, 4155). Attention is called to work on the same subject at the Shirley Inst.

C. J. WEST
Calculations and numerical data on the nitration of cellulose. L. SAUZAY. *Rev. gén. mat. plastiques* 7, 83-9, 139-45 (1931).—Indications are given regarding the control of nitrating acids and the prepn of mixts of predetd compn. A. P. C.

Nitration of cellulose with phosphoric-nitric mixed acids (preliminary communication). E. BRILL AND G. RUEFF. *Ber* 63B, 3212 (1930).—Nitration of cellulose with $\text{H}_2\text{PO}_4\text{-HNO}_3$ shows some marked contrasts with the ordinary $\text{H}_2\text{SO}_4\text{-HNO}_3$ nitration. Since H_2PO_4 does not sapon cellulose nitrates it is possible, with a suitably adjusted nitrating mixt (slight excess of P_2O_5), to obtain readily cellulose nitrates with 14.0-13.7% N. The nitration is accompanied by a marked swelling which facilitates the penetration of the HNO_3 into the interior of the fiber, resulting in a uniform nitration throughout. Unlike the product with much less N obtained with $\text{H}_2\text{SO}_4\text{-HNO}_3$, the cellulose nitrates prepd with $\text{H}_2\text{PO}_4\text{-HNO}_3$ are perfectly stable after very short boiling, they contain no trace of mixed $\text{H}_2\text{PO}_4\text{-HNO}_3$ esters producing instability. Nitrates with 11-11.5% N cannot be produced in fiber form with $\text{H}_2\text{PO}_4\text{-HNO}_3$. In the nitration with $\text{H}_2\text{PO}_4\text{-HNO}_3$ the lattice of the native cellulose changes into that of hydrate cellulose, whereas with $\text{H}_2\text{SO}_4\text{-HNO}_3$ not too rich in HNO_3 the lattice of the native cellulose is retained in the nitrate. With $\text{H}_2\text{PO}_4\text{-HNO}_3$ the nitration is relatively slow, and it can be observed that the HNO_3 converts the cellulose into mercerized cellulose before esterification begins. With $\text{H}_2\text{SO}_4\text{-HNO}_3$ contg. not too much HNO_3 , H_2SO_4 apparently enters the cellulose mol as ester quite rapidly, and the sulfate then reacts to a greater or lesser extent with the HNO_3 . Partially nitrated cellulose with the lattice of the native cellulose unchanged is no longer mercerized by $\text{H}_2\text{PO}_4\text{-HNO}_3$, i. e., the lattice of the native cellulose is retained in such nitrates. The nitrates obtained with $\text{H}_2\text{PO}_4\text{-HNO}_3$ give a clear point diagram of cellulose trinitrate when the nitration and washing are effected under tension. Details of the expts will be published elsewhere. C. A. R.

from a previously prep'd chart. II. Specific heat of Glauber's salt. KOTARO TANEMURA *Ibid* 324-5—The mol heat of H_2O of crystn of $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ was calcd as 96 cal. The mean mol heat of H_2O of crystn was detd to be 0.6-0.8 cal. The mol heat of $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ was calcd as 128.2 cal and the sp heat of the crystal at 25° as 0.40. III. Properties of sodium sulfate. KOTARO TANEMURA and SHIGEO MIYOSHI *Ibid* 330-1—A summary of the physicochem characteristics of Na_2SO_4 is given. H_2SO_4 lowers the transition temp and increases the soly of $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$; i. e., it lowers the crystn temp of the salt. MgSO_4 at first lowers the crystn temp of $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$, but increasing MgSO_4 concn beyond a certain limit increases the crystn temp, the solid phase then being $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$. With a normal viscose spinning bath for centrifugal systems, the salt which crystallizes out on cooling is always $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$, completely free from Na_2SO_4 and NaHSO_4 . It forms no double salts with ZnSO_4 , FeSO_4 , MnSO_4 , or CuSO_4 , which accumulate in the bath. IV. Dehydration of Glauber's salt. KOTARO TANEMURA, SHIGEO MIYOSHI and MIZUKI YOSHIDA *Ibid* 332-3—The vapor pressure of the system $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O} \rightleftharpoons \text{Na}_2\text{SO}_4 + 10\text{H}_2\text{O}$ at 25° is 18.4 mm Hg. The hydrous salt can be dehydrated in a closed chamber by circulating air dried with concd H_2SO_4 and NaOH . It also can be dehydrated by treating with twice its wt of 02% EtOH , or with concd caustic; 45.8 g with 100 g 50% NaOH soln yield 21 g Na_2SO_4 . Heating to the transition temp, centrifuging, cooling to 5° and centrifuging 3 times yields 31.05 g Na_2SO_4 from 100 g $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$. If 36.3% of the wt of the decahydrate is added as solid NaOH at the transition temp, the salt is dehydrated. V. KOTARO TANEMURA and SHIGEO MIYOSHI *Ibid* 333-4—In the method of sulfate recovery patented by the authors 0.9 kg Na_2SO_4 and 4 kg H_2O are assumed to accumulate as spent bath per 1 kg viscose rayon, 0.9 kg Na_2SO_4 combines with 1.1 kg H_2O on cooling. When 2.3 kg anhyd Na_2SO_4 is added to the spinning bath, and 7.3 kg of the decahydrate is removed on cooling, both 0.0 kg Na_2SO_4 and 4 kg H_2O are at the same time removed. The necessary equipment, crystallizer, hydroextractor and dehydrator, lasts much longer than an evaporator and is more efficient from the standpoint of heat utilization.

P. S. BILLINGTON

Manufacture of artificial silk. Spinning of viscose. II L. DE LERUW *Tech moderne* 22, 521-0, 623-7, 720-4 (1930), cf *C. A.* 24, 3643—A detailed description is given of the methods and app. used and methods of producing artificial silk other than viscose.

P. THOMASSET

Studies on lignin and related compounds. VII. A kinetic study of the action of hypochlorous acid on spruce lignin and its bearing on the constitution of the spruce lignin molecule. HAROLD HIOBERT and K. AUSTIN TAYLOR, *Can J. Research* 4, 240-53 (1931); cf *C. A.* 25, 2010—The action of HClO on glycol lignin in finely divided aq suspension, and in alk soln, is autocatalytic; the autocatalysis is due to formation of HCl , which catalyzes the reaction strongly. Apparently HNO_3 acts only by increasing the concn of undissocd. HCl . The effects of HCl may be a direct catalytic effect of undissocd HCl or may be due to formation of Cl_2 . A comparison of the amts. of HClO absorbed and HCl formed by methylated and unmethylated lignin indicates that if ketomethylene groups are present in lignin, they do not enolize sufficiently to play an appreciable part in the rapid reaction between HClO and lignin. The rapid absorption of HClO by lignin dissolved in alkali indicates that there is probably about 15% of a phenol nucleus in lignin. This value is supported by the rapid absorption by lignin, dissolved in alkali, of Br from Br water at very low concns and is based on the assumption of only 1 atom of halogen entering the phenol nucleus, which is true with vanillin. If 2 or 3 atoms enter the nucleus, this value of 15% must be divided by 2 or 3. It represents in any case a max value. The phenol nucleus accounts for only $1/3$, or less, of the total rapid HClO absorption. The remainder is probably due to addn of HClO to an ethylene linkage, chlorination of a phenol ether or a combination of both reactions.

A. PAPINEAU COUTURE

Alpha- and beta-lignosulfonic acids. PETER KLASON, *Svensk Pappers-Tid* 34, 118-9 (1931); cf *C. A.* 25, 1077—K gives the analytical data on which he based the formula $3\text{C}_6\text{H}_7\text{O}_2 + \text{H}_2\text{SO}_4 + \text{C}_6\text{H}_5\text{N} = \text{H}_2\text{O}$ (cf *C. A.* 24, 3358) and states that he has also succeeded in making the tetraconiferyl aldehydes.

WILHELM SEGERBLUM

Some notes on the recovery of soda. J. A. WALKER *Proc. Tech Sect Paper Makers' Assocn. Gl Brit Ireland* 11, Pt. 1, 186-202 (1930)—A good recovery of soda can be obtained only if an adequate proportion of black liquor is forthcoming at a strength and vol. which can be efficiently dealt with. The factors effecting this are discussed. The black liquor must be reduced to such a consistency as will permit a const combustion in the rotary. The factors on which the efficiency of the evaporator depend are discussed. Troubles which may arise in the system are described. The rotary is

discussed. The causticizing process is dealt with in some detail, and strict control is urged.

Soda process studies. III. Action of alkali solutions on sulfite pulp. J. H. ROSS and C. R. MITCHELL. *Research Notes* 3, 21 (1930). *Pulp Paper Mag Can* 31, 4172-3, 4144 (1931). Cf. C. A. 24, 4153. When bleached sulfite is treated in a rotary digester for 1 hr. with NaOH solns. at 150 MP, the concn. of alkali (over the range 1-7%) at temps. below 100° has but little effect on the quantity of material dissolved and on the α - and β -cellulose contents of the residue. At higher temps. the effect of concn. is much more marked. For every temp. there is an optimum concn. of alkali which produces a residue of max. alkali resistance and a concn. of alkali which produces a max. of bulking properties in the fiber. The fluidity of solns. of the treated pulp increases regularly as the temp. of digestion and alkali concn. are intensified.

Bleaching studies. I. An integrating attachment for the Hess-Ives tintometer. F. H. YORSTON. *Research Notes* 3, 11 (1930). *Pulp Paper Mag Can* 31, 3345-9, 3401 (1931). Cf. Ross, Mitchell and V. C. A. 25, 1078. The Hess-Ives tintometer has been supplied with a device for analyzing the tint of nearly white materials by multiple reflection of light at the surface of the material. By its use differences between the reflection factors of different bleached papers, measured through any particular color filter are magnified 2 fold. Further, since the magnification increases with the value of the reflection factor, differences in hue are greatly accentuated. **II. Measurement of hydrogen-ion concentration of bleach liquors with the glass electrode.** *Research Notes* 3, 15-16 (1930). *Pulp Paper Mag Can* 31, 374 (1931). -Consistent and accurately reproducible measurements of the H-ion concn. of bleach liquors can be made with the glass electrode and a vacuum tube potentiometer. HOCI acts as a buffer in bleach liquors within the range of pH 6.5-8.5.

Some notes on the Southern yellow pine district—the large sulfate wood reserve in the United States. FOLKE JOHANSSON. *Svensk Papperstidn.* 34, 82-7 (1931). -Descriptive.

The composition of Diospyros ebenum Koen. YOSHIZO SHINDO. *Cellulose Ind (Tokyo)* 6, Abstracts 37 (1930). -The sample analyzed as follows: moisture 10.3, benzene ext. 15.1 sol. in 0.2% NaOH 4.5, lignin 30.8, Cross and Bevan cellulose 30.9, purity of cellulose 99.5, pentosans 13.8%. Xylose, mannose and galactose were identified.

Papermaking properties of phormium tenax (New Zealand Flax). MEALE R. SHAW, GEORGE W. BICKING and MARTIN J. O'LEARY. *Bur Standards J. Research* 6, 411-20 (1931). -The caustic soda process and 2 stage cooks using sodium sulfite and caustic soda resp. produced good results as to quality and yield for wrapping and writing papers.

Utilization of industrial by-products with particular reference to the pulp industry of the United States. ROBERT W. GRIFFIN. *J. Am. Leather Chem. Assoc.* 26, 180-6 (1931). -An address.

Wood-pulp flour and shives in mechanical wood pulp. O. M. HALSE. *Proc. Tech. Sect. Paper Makers' Assocn. Gt. Brit. Ireland* 11, Pt. 1, 18-20 (1930). -See C. A. 24, 5488.

The Norwegian pulp and paper research laboratory, 1930. MAGNUS CHRISTENSEN. *Proc. Tech. Sect. Paper Makers' Assocn. Gt. Brit. Ireland* 11, Pt. 1, 3-17 (1930). -Descriptive.

Estimation of wood pulp in paper. E. DEBENEDETTI. *Boll. staz. sper. carta e fibre tessili vegetali* 9, 134-7 (1930). -The chem. methods for detecting and detg. wood pulp in paper utilize reactions for cellulose and lignin. The lignin content varies with different wood pulps, but av. values have been detd. for the woods and celluloses used in papermaking, and from these values can be calcd. approx. the wood or mechanical pulp contents in a paper when jute and manilla are absent. Reagents used are: (a) 1 g. phloroglucinol, 500 cc. alc., 25 cc. concd. HCl; (b) 5 g. aniline sulfate, 100 cc. dist. water, 1 drop H₂SO₄; (c) dimethyl *p*-phenylenediamine. Their use is limited to white or light-colored papers. The quantity of lignin is detd. (1) directly by liberating from the cellulose and weighing, (2) by measuring the consumption of certain chemicals absorbed by the lignin substances, (3) by dissolving the lignin with suitable agents, weighing the residual cellulose, and calcg. the lignin by difference. Many methods use the hydrolysis of cellulose with strong acids that under proper conditions do not attack the lignin, which can be dried and weighed. H₂SO₄, HCl and HNO₃ alone, or together or with dehydrating agents, have been proposed. Cl, Br, phloroglucinol and KMnO₄ also are used for detg. lignin.

Pulping with chlorine. I. Pulping bagasse. J. KAWAMURA. *J. Soc. Chem. Ind.*

Japan 33, Suppl. binding 502-5(1930)—Air-dry bagasse contained moisture 12.55, Cella ext. 0.94, alc. ext. 1.09, H_2O ext. 2.84, 1% NaOH (hot) ext. 29.0%. The NaOH-extd. residue contained ash 3.5, lignin 22.3, cellulose (Cross and Bevan) 43.4, pentosans 27.4%. The cellulose contained 69.9% α -cellulose and 9.3% pentosans. Bagasse after extn. with 0.5, 1.0 and 2.0% NaOH was chlorinated at about 50% consistency, the change in yield, lignin, pentosan and ash being followed. The more drastic the alk. pre-treatment, the lower were the yield and lignin content for an equal Cl consumption. The pentosan content remained fairly const. at 31-33% of the chlorinated material. Of this, $\frac{1}{4}$ was removed by treating for 12 hrs. with cold 10% NaOH, 7% cellulose being likewise removed. II. Reactions of chlorine. 1. *Ibid* 390-8.—The ratio of Cl as HCl to organically combined Cl formed in chlorinating ligno-celluloses has been variously reported from 1.13 to 1.04. K. extd. bamboo with alc. benzene and 1% NaOH, chlorinated, washed and extd. the residue 3 times with 1% NaOH. HCl was detd. in the wash water, after expelling free Cl, while organically combined Cl was detd. by heating the NaOH ext. at 250° and detg. as NaCl. The ratio of acid to organically combined Cl was 1.04. III. Reactions of chlorine. 2. *Ibid* 34, Suppl. binding 62-4(1931)—Results similar to those reported above for bamboo were obtained with bagasse. When a large excess of Cl was allowed to act on the bagasse for a long time, the ratio of HCl to org. chlorides was slightly increased, indicating an oxidation.

M. HERNIG

The technic of purifying pulp: effect of calcium lime in the production of a bleached sulfite with a high α -cellulose content. G. P. GENBRG *Pulp Paper Mag. Can.* 31, 19-210, 219(1931)—By cooking bleached or partly bleached sulfite pulp with lime, the α -cellulose content of the final residue is increased appreciably, accompanied by a loss in wt. depending principally on the original α -cellulose content. More than 5-10% CaO has no beneficial effect on the final results and on the loss in wt. of the pulp. More than 70% of the total increase in α -content is effected in the first 4 hrs. at the consistency (4.5%) and temp. (100-140°) used. The most important factor is the initial α -cellulose content, and in order to obtain under most economical conditions, a final pulp with a max. α -content, it is necessary to start with a pulp so treated that a max. α -content is retained before the CaO treatment is begun. A multiple treatment with CaO has no increased effect on the α -cellulose content, but produces addnl. loss in wt. By omitting washing after bleaching and by starting the CaO treatment immediately after exhaustion of the bleach, a higher α -cellulose content is obtained. Practically the same α -cellulose content is obtained by treatment with 10% CaO and with 3% NaOH, but the CaO-treated pulp has a lower Cu no. and suffers a higher loss in wt. The optimum conditions for max. α -cellulose, min. bleach consumption and min. shrinkage were: Unbleached sulfite is given a preliminary bleaching with a ratio of 35% bleach powder to Roe Cl no. of 2:1, treated with CaO, washed and given a final bleaching to a full white color with a 0.75:1 bleach Cl no. ratio. If the ratio of bleach to Cl no. is controlled at 2:0.1, the final α -cellulose content will be independent of the Cl no. of the unbleached pulp, but the shrinkage will be higher with increased Cl no. The use of MgO lime instead of CaO is of no advantage. The time to exhaust the bleach is prolonged, and a lower α -cellulose content is obtained when the CaO is added to the pulp at the same time as the bleach liquor and when the bleaching action is allowed to take place in the presence of the CaO. Corn trial runs made by the procedure developed in the lab. produced a sulfite pulp with high α -cellulose content. Sulfate pulp is not affected to any extent by either CaO or NaOH under the conditions applied to sulfite pulp in the investigation. A. P.-C.

The Rosen process for the production of sulfite pulp from resinous woods. E. MEUNIER *Papeterie* 33, 214-8(1931)—A brief discussion of the process and its merits. The difficulties encountered in pulping resinous woods by the sulfite process have been shown to be due to the fact that the compn. of the resin acids in the heartwood apparently differs from that of the resin acids in the sapwood. Rosen overcomes the difficulty by a preliminary treatment of the chips in the sulfite digester with a weak alk.-earth soln. for 2-5 hrs. at a temp. of up to 110°, followed by washing with salt soln.; this treatment can be repeated if necessary; the chips are then treated with SO_2 gas, and are then subjected to the usual low-temp. sulfite cooking, yielding a very pure pulp, low in ash and contg. up to 92.4% of α -cellulose, which can be bleached with 2.5-4% available Cl.

A. PAPINEAU-COUTURE

Determining the bleachability of pulps. R. KARLBERG *Paper Trade J.* 92, No. 11, 41-4(1931)—A review is given of methods developed for the detn. of the amt. of bleaching powder or Cl required for bleaching pulps, in the form of abstracts taken mostly from the German literature.

A. PAPINEAU-COUTURE

The childhood of paper making, as illustrated by Kashmiri methods of the present

day W. RAITH *Proc Tech Sect Paper Makers' Assocn Gt Brit Ireland* 11, Pt 1, 21 (1943) cf *C A* 24, 1975.—Historical. Twenty five photographs of great interest to the modern paper maker are given. P. S. BILLINGTON

Fundamental research and the paper industry. T. TREVOR POTTS *Proc Tech Sect Paper Makers' Assocn Gt Brit Ireland* 11, Pt 1, 115-20 (1930).—A résumé is given of past fundamental investigations and the suggestions of some problems for future study. P. S. BILLINGTON

What patents mean to the paper technologist. JOSEPH ROSSMAN *Paper Trade J.* 92, No 14 58-60 (1931). R describes the classification system of the U. S. Pat. Office, defines the patent holders' rights, interprets the validity of claims, discusses the sulfate process patent of Eaton and Dahl and shows the advantages of patent protection. A. PAPINEAU COUTURE

The development of paper manufacture from southern pines. R. H. STEVENS *Paper Ind* 12, 2019-20 (1931) cf *C A* 25, 590.—A general discussion of the advantages and possibilities of the Southern U. S. for the production of pulp, particularly kraft, and paper. A. PAPINEAU COUTURE

The formation of paper. W. HOWARD CAMPBELL *Research Notes* 3, 1-3 (1930); *Pulp Paper Mag Can* 31, 424-5, 438 (1931).—An address dealing particularly with the qualities of pulp which make it suitable for paper. The adhesion of cellulose to cellulose is particularly stressed, and C's theories as to the mechanism of the so-called hydration by heating (*C A* 24, 1976) are explained. A. PAPINEAU COUTURE

Corrosion reduced by new constructional materials. JAMES A. LEE *Paper Trade J.* 92, No 12, 67-8 (1931).—A discussion of the merits and possibilities of alloy steels (particularly 18-8 alloy and others of the same type) and of C in the sulfite and paper industries. A. PAPINEAU COUTURE

Simple electrometric methods of analysis in the paper mill laboratory. T. TREVOR POTTS *World's Paper Trade Rev* 95, 1037-8 (1931).—Callan and Horrobin's app. (*C A* 23, 789) did not give good results for alkalimetric work with the circuit in the original form suggested by C and H, inserting a potentiometer across the output of the transformer allows of much better control of the voltage of the input to the electrodes, enables better adjustment of the initial reading of the measuring instrument, and prevents damage to the latter if the soln. under examn. happens to be of high cond. The outstanding use of this method in the paper mill lab. is the examn. of highly colored caustic liquors, but it is also applicable to other titrations, such as sulfates with BaCl₂ and pptn. of size by alum. Attention is drawn to the effects of the presence of dissolved CO₂ in ordinary distd. H₂O on the potentiometric detn. of H⁺ ion concn. For ordinary paper mill work, direct detn. on the stock taken from the mixing box or stuff chest is preferable to extn. of paper with distd. water. The quinhydrone electrode used in conjunction with the said calomel electrode is by far the most useful in the paper mill lab. A. PAPINEAU COUTURE

Notable recent improvements to high-speed paper-making machines. S. RICHARDSON *Proc Tech Sect Paper Makers' Assocn Gt Brit Ireland* 11, Pt 1, 131-43 (1930).—The increased production from paper machines has been due to increased paper speed, increased wire width, and reduction of over all mfg. costs. These 3 factors are discussed in some detail. P. S. BILLINGTON

The application of electrical power in paper mills. JAMES R. HAPPER *Proc Tech Sect Paper Makers' Assocn Gt Brit Ireland* 11, Pt 1, 164-79 (1930).—The a. c. and d. c. systems are compared. The placing of a heater-drive motor inside the beater-roll has worked out very satisfactorily. The merits of steam vs. electricity are discussed. P. S. BILLINGTON

The testing of condenser paper. F. L. ROMAN *Proc Am Soc Testing Materials* 30, Pt. 11, 1012-24 (1930).—A discussion is given of the relationship between several characteristics of condenser paper made from linen stock and the dielec. strength (break-down voltage) of the condensers wound with this paper. Variations in the density of the condenser paper have no noticeable effect on the dielec. strength of the corresponding condensers. It is shown that, when small nos., conducting paths through the paper have no important effect on the dielec. strength of condensers wound with 2 sheets of the paper between tin foils, but the percentage of "shorted" condensers or of condensers failing at very low voltages may be expected to increase approx. as the square of the no. of conducting paths per unit area of the paper. Variations in the porosity of the paper are shown to have marked influence on the dielec. strength of the corresponding condensers and porosity requirements are suggested for papers used in low voltage condensers and for papers used in high voltage condensers. P. S. BILLINGTON

Variability in results met with in paper strength testing. CYRIL V. OLIVER *Proc*

Tech Sect Paper Makers' Assoc. Gt Brit. Ireland 11, Pt. I, 53-84(1930) —See C. A. 24, 3366 P. S. BILLINGTON

A critical study of the bursting-strength test for paper. F. T. CARSON AND F. V. WORTHINGTON *Bur Standards J Research* 6, 339-53(1931) —See C. A. 24, 1977.

P. S. BILLINGTON

Correlation of bursting strength and tensile strength of paper. JAMES STRACHAN. *Proc Tech Sect Paper Makers' Assoc. Gt Brit Ireland* 11, Pt. I, 80-114(1930), cf. C. A. 24, 5485 —A math treatment of the subject, and a discussion of the formulas obtained by various investigators. The methods of calen indicated give a means of investigating the degree of cohesion produced between the fibers by papermaking operations, as reflected by the bursting test. A voluminous discussion is included. P. S. B.

The coloring of paper. N. L. MATHEWS *Proc Tech Sect Paper Makers' Assoc. Gt Brit Ireland* 11, Pt. I, 144-63(1930), cf. C. A. 24, 5485 —A description of the nature and source of natural vegetable dyes and pigments is given. The theory of the org constitution as related to color is briefly discussed, and the theory of dyeing is touched upon. The characteristics of direct or substantive, basic, and vat dyestuffs are discussed. The cause of two sidedness is explained, and the reason for adulteration of the dyestuffs is discussed. P. S. BILLINGTON

Classes of color used in the dyeing of paper. HAROLD E. PRATT *Paper Trade J* 92, No. 14, 57(1931) —A brief review of the application of basic, acid and direct dyes and pigments to paper coloring. Their limitations are pointed out. A. P. C.

Further study of paper-coating minerals and adhesives. MYRL B. SHAW, GEORGE W. BICKING AND MARTIN J. O'LEARY *Bur Standards J Research* 5, 1189-1203(1930), *Paper Trade J* 92, No. 14, 52-7(1931), cf. Hamill, Gottschalk and Bickling, C. A. 19, 175, 20, 1510, 21, 175 —American clays examd compared favorably with foreign clays as paper-coating minerals and produced good printing qualities. Diatomaceous earth compd did not adhere well to the base paper. The modified starches used did not appear to have as strong adhesive qualities as casein or glue. Graded according to their clay suspending property, the adhesives were in the following descending order: starch, casein, glue. Chem analyses, color measurements and settling data for the minerals used are given. A. PAPINEAU-COUTURE

Animal and vegetable adhesives for paper manufacture. GIULIO C. CIOCCHI *Ind cartia e artis grafiche* 33, 265-8, 512-3(1930) —The manuf of glue from bones, hides, etc is described. The material generally employed in paper manuf is obtained from the fleshy side of hides, butcher's residues, hide cuttings, and calf and mutton heads. Animal glue is a positive colloid, its solns represent dispersed systems with a very fine phase. The reaction of the gelatin must be nearly neutral, since the presence of acids reduces and hinders its fixation in paper. Microorganisms transform gelatin into an unfixable product. SO_2 eliminates this danger. A good glue absorbs 10-12 parts of water. The value of glue depends on the water and ash contents, m. p., viscosity, swelling capacity, fatty matter, acidity, mineral impurities, the elastic force and adhesive power, all of which are influenced by the p_H . Gelatin has two max swelling points in water, $p_H = 2.4$ and $p_H = 11.6$, and has a min swelling point at $p_H = 7.7$. In the hydrolysis of ossein there is a min between $p_H 4.7$ and $p_H 6$, the reaction proceeds well with formation of a good product at $p_H = 3-4$. The decompn. of gelatin into proteoses and peptones follows in a 2nd hydrolysis, it is not strong between $p_H = 3$ and $p_H = 8$, but takes place beyond this. To avoid hydrolysis during the evapn, the soln is brought to $p_H = 4.7$. Casein must be produced under conditions that effect coagulation in an acid medium with $p_H = 4-4.1$. At this hydrogen ion concn is obtained a product contg. the min quantity of salts. The greatest yield with starch pastes is below 100° , above that temp. alteration follows. Stacey gum is starch solubilized by oxidation; Osalon and Vosgeline are farina treated with O_2 . Dextrin prepd. with farina gives a limpud water soln not obtained with starch from cereals. The cohesive power of dextrins decreases with the increase of their soly. in cold water. Gums are not well defined chem products. Kordofan gum destroys the capillarity of water solns, having the same viscosity in cold and warm solns. Gum arabic solns of 20°Bé are used in paper making. Gum Senegal is less sol. in water than gum arabic, and its more viscous and gelatinous solns sep. the gum more easily with many chemicals. R. SANSONE

Finishing, processing and converting of paper. NORMAN CLARK *Paper Trade J* 92, No. 13, 53-5(1931) —After a brief summary of developments in the field of finishing, processing and converting paper during the year 1930, the Tech. Assoc. of the Pulp and Paper Industry literature relating to coated and processed papers is reviewed. A. PAPINEAU-COUTURE

The porosity of paper. GIRO SILVIO *Papier* 34, 173-81(1931); cf. C. A. 24,

6011 The effects of the thickness of the sheet and the pressure on the detn. of the porosity are studied. The tests were carried out on 100% soda, unloaded and uncalendered paper, all the samples being made from the same beaterful of stock to eliminate possible differences in degree of beating and sizing. *Effect of thickness*: (1) The time required for the passage of a given vol. of air through the sheet does not vary directly with the thickness, at least above a certain thickness, which might be called the limiting thickness. (2) Above the limiting thickness, the time is practically const. for a given pressure. *Effect of pressure*: By defining the impermeability as the reciprocal of the porosity and calcg. it from the formula $pt/2$, in which p is the pressure and t the time, it was found to be an essentially variable property, increasing with the pressure at which it is detd. but varying only slightly above a certain pressure. As it has been suggested that the impermeability is rather a function of \sqrt{p} , it was calcd. from the formula $t \propto f(\sqrt{p})$ but was found to be even more variable when expressed in this form. An attempt to find a formula which would give a practically const. value for the impermeability at different pressures led to the formula $t = t/2 \times p^{1/2}$, but even this did not give absolutely const. values. It is therefore considered preferable always to carry out the detn. at a definite pressure, which should be selected either in the zone where small variations in p lead to small errors in t , or in the zone of utilization of the paper.

A. PATTINAU-COUTURE

Stainless steels used in heavy machining [in paper mills] (Head) 9. Animal materials for artificial silk (Meydel) 29. Structure of the crystal lattice of cellulose (Asbury, Marwick) 2. Swelling of cellulose in HClO_4 (Andrew, Reinhardt) 2. Cellulobiosan and cellulose (Hess, Gartner) 2. Bleaching machines (Ger. pat. 518,613) 25. Apparatus for clarifying waste waters of the paper, cellulose, etc., industries (Ger. pat. 518,063) 14. Resins from cellulose material (Russ. appl. 50,420) 26. Ultramarine [from sulfite liquor] (Russ. appl. 50,674) 26.

LUHMANN, F., AND ESSLINGER, R. Die Fabrikation der Dachpappe und der Anstrichmasse für Pappdächer. 3rd ed. Leipzig: A. Hartleben's Verlag, 172 pp. M. 4, bound, M. 5.

TROTMAN, S. R., AND TROTMAN, E. R. Artificial Silks. London: Griffin, 274 pp. 18s., net.

Cellulose. I. G. FARBERMUND A-G (Heinrich Diekmann and Christoph Beck, inventors). Ger. 518,154, Jan. 9, 1929. In the manu. of cellulose by treating wood with HNO_3 , the concn. of the acid is maintained during the treatment by leading NO_2 or gases contg. NO_2 into the reaction mixt.

Cellulose. ZELLSTOFFFABRIK WOLOKOF, VALENTIN HOTTENROTH and OTTO FAUST. Fr. 37,121, June 22, 1929. Addn. to 663,684 (C. A. 24, 722). Cellulose of all kinds is improved by a treatment with NaOH , the temp. of treatment and the concn. of the NaOH being adapted to one another, the lower the concn. of the soln., the lower being the temp. used.

Treatment of grasses for the production of cellulose. L. OADÓDY, B. OADÓDY, Z. HOLLÓSY and E. HOVÁTH. Hung. 101,627, May 12, 1930. Pectic substances and lignins are decompd. by boiling under pressure with sufficient Ca(OH)_2 to neutralize the acid decompn. products. The silica of the raw material may be extd. by hydrides, carbonates or sulfides of alkali metals. Cf. C. A. 24, 5495.

Yellow pigmented cellulose. EMUL DE STUBNEA. U. S. 1,705,764, March 10. Various details are described for pptg. a yellow pigment from Pb acetate, $\text{K}_2\text{Cr}_2\text{O}_7$ and H_2SO_4 solns. on "sol. fibrous cellulose" such as cellulose ester or ether material in such finely subdivided form that when the material is dissolved, the pigment becomes dispersed in the soln. without subjection to grinding or other special dispersing treatments.

Cellulose esters. BRITISH CELANESE, LTD. Brit. 338,745, Dec. 11, 1928. Cellulose material such as cotton, cotton linters, reconstituted cellulose or wood pulp is subjected to successive pretreatments with HNO_3 (for an hr. at about 50°) and an aliphatic acid anhydride such as Ac_2O or propionic or butyric anhydrides (for an hr. at room temp.) and may then be esterified by addn. of a catalyst such as H_2SO_4 or ZnCl_2 and, if desired, a further quantity of the anhydride. Products suitable for making rayon, films, etc., are obtained.

Cellulose esters. I. G. FARBERMUND A-G. Brit. 338,798, Feb. 25, 1929. See Fr. 687,627 (C. A. 23, 813).

Apparatus for digestion of nitrocellulose or for purification of cellulose. M. G.

MILLIKEN (to Hercules Powder Co.) Brit. 338,941, May 29, 1929. In an app. of the character described in Brit. 301,267 (C. A. 23, 4971), the material is forced under pressure together with water through a restricted passage such as a coil (which may be formed of chrome steel of small diam. and a length of 2500–3000 ft.) which leads directly or through a suitable valve to a stand pipe or chamber to effect substantial equalization of pressure in the coil between the inlet and discharge ends. The coil may be heated by steam to 130–160°. Various details of app. and operation are described.

Apparatus for saponifying cellulose acetate. VERGIN FOR CHEMISCHE INDUSTRIE A.-G. (Karl Werner, inventor). Ger. 515,779, Aug. 12, 1928.

Cement for films. M. SARY. Hung. 100,689, Oct. 3, 1929. Ten g. celluloid is dissolved in 5000 g. acetone and 500 g. amyl acetate, and 10 g. AcOEt is added.

Bleaching tower for fibers, especially cellulose. PAPER PATENTS CO. Ger. 515,961, Nov. 30, 1928.

Bleaching unspun fibers. PAPER PATENTS CO. Ger. 515,731, Aug. 5, 1928. The bleaching of raw material, especially cellulose, for paper manuf. is described.

Modifying the luster of rayon at the time of production. H. DREYFUS. Brit. 338,490, May 17, 1929. In making filaments or other products from various specified cellulose esters or ethers, a reduced or modified luster is obtained by adding to the solns. used for dry or wet spinning a small proportion (suitably 0.5–2%) of finely subdivided insol. materials (preferably white pigments) such as BaSO_4 , Pb sulfate, ZnO , Ba phosphate, ZnCO_3 , CaCO_3 , MgCO_3 , CuSO_4 , Al_2O_3 , or silica, and, to minimize settling, materials of low sp. gr. may be added such as Li and Mg silicates, oils, dispersing agents, etc., also may be used. The use of alkyl and aralkyl ethers of starch also is mentioned, and various details and modifications of procedure are described.

Artificial silk from viscose. I. G. FARBENFABR. A.-G. (Adolf Kämpf and Arnulf Hager, inventors). Ger. 518,234, Apr. 27, 1929. Artificial silk of reduced luster is prep'd by spinning viscose solns. to which a small quantity of a stannate or antimonate, particularly a pyroantimonate, has been added. The added salt is decompd. during the spinning process, with the result that colloidal metal oxide is deposited in and on the fibers. An example is given.

Artificial silk, etc. LÉON LILJENFELD. Fr. 37,344, July 31, 1929. Addn. to 666,178 (C. A. 24, 1509). Artificial threads of a high resistance in the dry state, above 2 g. per denier, are made by coagulating a thread of viscose by means of a gaseous mineral acid.

Artificial silk spinning. HANS EGGERT. Ger. 515,871, July 14, 1929. The adhering pptg. bath is removed from spun artificial silk fibers by suction-washing the fibers after they are wound on spools.

Artificial filaments, ribbons, etc. BRITISH CELANESE, LTD., R. P. ROBERTS and L. W. GREGORY. Brit. 338,936, July 29, 1929. Opaque products formed from org. esters or ethers of cellulose coagulated with aq. media are given improved transparency and tensile strength by heating them (suitably with dry steam or other hot gases or vapors while under tension, and if desired after preliminary impregnation with solvents, swelling agents or plasticizers). Various details of procedure are described.

Centrifugal spinning apparatus for artificial silk manufacture. W. S. MUMFORD. Brit. 338,958, Aug. 29, 1929. Structural features.

Piston pump for solutions of viscose, cellulose acetate, etc. TEXTILES ARTIFICIELS ING. A. MAURER. Ger. 518,100, June 22, 1929.

Spinning pumps for artificial silk. LOUIS CLAUDE BRUN. Fr. 37,287, July 30, 1929. Addn. to 658,463 (C. A. 23, 5317).

Apparatus for dry-spinning of artificial silk filaments. NELSONS SILK, LTD., and J. NELSON. Brit. 338,881, July 25, 1929. Structural features.

Pulp from wood. LINN BRADLEY and EDWARD P. MCKEEFE (to Bradley-McKee Corp.). U. S. 1,795,754, March 10. Wood is digested in a cooking liquor contg. most of its soda in the form of a sulfite of Na; the residual liquor is subjected to a treatment including a reducing furnacing operation to yield a product contg. Na_2CO_3 and a relatively large amt. of Na_2S ; and the Na_2S thus formed is utilized in an alk. cooking liquor to cook an addnl. quantity of wood. Residual liquor from this second cooking operation is subjected to a furnacing to yield a product contg. Na_2CO_3 and a lesser proportion of Na_2S than was present in the alk. cooking liquor, and the furnace product is subjected to a sulfiting to produce a cooking liquor contg. a quantity of Na_2SO_3 at least equiv. to any $\text{Na}_2\text{S}_2\text{O}_3$ present; and the cycle is repeated, by use of the last mentioned cooking liquor in the first stage of the process. U. S. 1,795,755 relates to a cyclic process, in which one lot of wood is cooked in a liquor contg. SO_2 and NaHSO_3 , and another lot of wood is cooked in a liquor contg. NaOH and a lesser quantity of Na_2S , residual liquors from these

2 operations being mixed and suspended matter sep'd. This is followed by addn of a Na comp'd such as NaOH capable of redissolving remaining suspended org. material, concn. treatment of a portion of the conc'd liquor by a reducing furnacing operation to form Na_2CO_3 and a considerable amt. of Na₂S, treatment of another portion of the liquor by furnacing to produce mainly Na_2CO_3 without much Na₂S, forming sep. liquors from these 2 furnace products, causticizing the liquor contg. the greater proportion of Na₂S to form a cooking liquor contg. NaOH and Na₂S, and subjecting the other liquor to treatment to form a cooking liquor contg. NaHSO_3 and SO_2 and not contg. any objectionable quantity of $\text{Na}_2\text{S}_2\text{O}_3$. U. S. 1,795,756 relates to a generally similar process in which however the wood cooked in alkali is treated with a larger proportion of Na₂S than NaOH when freshly charged into the digester. Cf. C. A. 25, 816.

Pulp from material such as wood chips. EAWIN H. HUSSEY (to Bauer Bros. Co.) U. S. 1,795,603, March 10. Material such as wood chips is disintegrated in the presence of sufficient water to thoroughly impregnate the material but without any substantial excess of water, further reduction and disintegration is effected (suitably by rotating disks) in the presence of a large quantity of water, followed by rolling, pressing and squeezing of the material, without appreciable cutting, between rotating, relatively movable disks having teeth inclined in an opposite direction to the direction of rotation. App. is described.

Bleaching wood pulp. LLOYD BRADLEY and EDWARD P. MCKEEFE (to Bradley-McKeeffe Corp.) U. S. 1,795,767, March 10. Fibrous material from an incomplete digestion of wood is treated with liquor comprising a manganate such as NaMnO_4 together with an alkali metal hypochlorite. U. S. 1,795,758 describes numerous details and modifications of procedure of a process which may comprise digesting raw cellulosic materials in successive batches, one in an alk. cooking liquor and the other in an acid cooking liquor, and forming each liquor with inorg. comp'ds. derived and recovered from the other liquor after the cooking operation therewith. Cf. C. A. 24, 4634.

Paper pulp. MEAD PULP & PAPER CO. Brit. 339,333, Aug. 28, 1929. In a continuous process of pulp production, stock of high consistency is treated with a gaseous chlorinating agent as it is fed down a tower (which may be fed with gaseous Cl from a liquid Cl supply). Numerous details of arrangement and construction of app. are described. Cf. C. A. 24, 4392.

Paper pulp. MEAD PULP & PAPER CO. Brit. 339,334, Aug. 28, 1929. For producing a free pulp of high strength, color, and clay-carrying quality suitable for paper making, from wood, straw, fibrous grasses and the like, such as that which has been subjected to an incomplete chem. digestion (e. g., leached chestnut chips), the material is subjected to multi stage and coordinated chlorinating treatment, the first stage being with an insufficient quantity of chlorinating agent acting at a limited temp. and the second stage involving treatment of a thickened pulp with a stronger chlorinating agent. Numerous details of app. and procedure and auxiliary and preliminary treatments of the material are described.

Rotating perforated drum apparatus for separating paper pulp, peat fiber, etc., from liquids. E. F. SUNDBSTRÖM and K. E. STEEN. Brit. 338,547, July 18, 1929. Structural features.

Treating black liquor from the soda-pulp process. ERIK HÄGGLUND. Ger. 518,062, June 24, 1926. See Brit. 273,267 (C. A. 22, 1854).

Treating black liquor formed in the soda-pulp process. ERIK HÄGGLUND. U. S. 1,795,557, March 10. The liquor is subjected to a heat and pressure treatment to ppt. most of the org. substances present as carbonaceous and pitchy material. After sep'n of the liquor from the ppt. thus formed, it is causticized (the acetate content being left substantially unchanged) and conc'd, and a portion of the NaOAc content is crystd. out. Another portion of the causticized liquor is used to increase the alkalinity of the black liquor, and a further portion is used to act upon pulp-forming fibrous material, crystd. NaOAc and volatile products are recovered. An arrangement of app. is described. Cf. C. A. 21, 3128, 24, 4932.

Paper. FELIX THUNERT. Ger. 518,099, Dec. 24, 1924. A method is described of making packing, etc., paper having a downy upper surface, by securing cellulose wadding to a smooth paper support by means of an adhesive.

Paper, molded sheets, blocks, etc., from fibrous materials. RESPATS, INC. Brit. 339,528, Jan. 11, 1930. See U. S. 1,770,430 (C. A. 24, 4634).

Bituminous paper board. A. SOMMER HUNG. 100,522, Aug. 7, 1929. Molten asphalt is emulsified with paper pulp by mixing and heating. The emulsion is broken down by drying, and paper board is made from the solid phase by rolling.

Roll for calendaring apparatus. THEODOR GREIS. U. S. 1,795,558, March 10. A

calendering roll comprises a core around which is wound a band or ribbon of material standing on edge and having its web threads arranged in staggered relation.

Imparting high finish to paper. ALBERT L. CLAPP (to Bennett, Inc.). U. S. 1,797,048, March 17. After the drying operation, there is applied to paper an aqueous dispersion of solid particles of waxes flexible at the calendering temp. of the usual hot calendering rolls at the dry end of the paper machine, and the paper is hot calendered. App. is described.

Apparatus (with twin glazing cylinders) for glazing paper. J. W. GRANT. Brit. 339,370, Sept. 13, 1929. Structural features.

Waterproof flexible paper. THE SCUTAN CO. Ger. 514,922, Nov. 25, 1927. See Brit. 281,316 (C. A. 22, 3530).

24—EXPLOSIVES AND EXPLOSIONS

CHARLES E. MUNROE AND C. O. STORM

Fifty years' experimental research upon the influence of steam on combustion of carbonic oxide. WM. A. BONE. *J. Chem. Soc.* 1931, 338-61.—An historical account is given of expts. and theories on the combustion and explosion of gaseous mixts. from the time of H. B. Dixon's discovery of the activation by water vapor of the combustion of CO (1877) up to now. Chem. and phys. theories of gaseous explosions are considered in separate sections.

The laws of combustion of colloidal powders containing vaseline. H. MURAUER AND G. AUNIS. *Compt. rend.* 192, 418-21 (1931). cf. C. A. 24, 4396.—The variation in $spdt$, in which p is the pressure prevailing during combustion and t is time, was studied in powders contg. (1) vaseline and (2) centralite ((1-PhN)₃CO) to lower the temp. of explosion of the nitrocellulose nitroglycerin mixt. The decrease in $spdt$ with d of charge is much less with vaseline than with centralite. The powders contg. vaseline obey the same laws of combustion as other powders at high d s of charge but not at d s corresponding to pressures below 500 kg. per sq. cm. The peculiar behavior of vaseline is attributed to the fact that it is emulsified rather than dissolved in the powder.

F. L. BROWNE

Dust explosions, with special reference to wood-working industries. H. R. BROWN. *Am. Soc. Mech. Eng.*, Advance paper, Dec. 1-5, 1930, 3 pp.—Explosions in flour mills have been common, but now this hazard is recognized in grain elevators, starch factories, sugar refineries, wood-working plants, textile mills and other factories in which materials in form of dust are manufd. Precautionary measures are suggested and regulations have been tentatively adopted for prevention of dust explosions in wood flour mfg. establishments.

E. I. S.

The constitution of atmospheres behind stoppings. T. DAVIO JONES. *Colliery Guardian* 142, 398-401, 493-4 (1931).—Many analyses are reported of mine air contained in sections of coal mines which, for one reason or another, have been sealed off. Four different types are distinguished which are characteristic of the action taking place in the distinct sealed off. (1) The air is almost unchanged with time. There is little formation of CO₂. (2) The air is more or less rapidly replaced by CH₄, the CO₂ formed from the original air being pushed out. In one case CH₄ reached 90% in 12 weeks, the addnl. 10% being N₂. (3) The air is replaced by a mixt. of CO₂ and N₂ (black damp). In one unusual case black damp (with 18% CO₂) reached 100% in 134 weeks. This type is due to fires. When the O₂ has been reduced below 10% flames cannot exist, although heating may continue until the O₂ has been used up. (4) This results from a race between the accumulation of fire damp and the diminution of O₂. From a safety point of view explosions do not occur when the CH₄ exceeds 15% or the O₂ is less than 10%. Generally the CH₄ wins the race. Examples of consecutive analyses and diagrams illustrating the 4 types are given.

A. L. KIBLER

Role of gas masks in mines (Esztró). 13. The propagation of flame in electric fields (GUÉNAULT, WHEELER). 2. Use of HCl to prepare bore holes for blasting (Brit. pat. 338,454) 13.

Explosive derived from hexamethylenetetramine. E. L. D'ASTRECK. Brit. 339,024, Sept. 20, 1929. (CH₂)₆N₄ is oxidized with H₂O₂, the resulting product is treated with HNO₃ and is further oxidized with H₂O₂. Various details of procedure are described. Blasting cap. WM. PRITCHARD, JR. (to Hercules Powder Co.). U. S. 1,795,440,

March 10 A weighting material such as Pb is placed in a cap comprising a caving and a charge which may be formed of fulminate-chlorate compn. in order to give the cap as an entirety a sp gr higher than that of coal, so that mech. sepn from the broken coal is facilitated

Diffusion apparatus for indicating the presence of combustible gases in air. BRUNO TREBITSCH Austrian 120 956, Aug 15, 1930

Testing mixtures of coal dust and stone dust in mines. EMIL WITTE Ger. 472,529, Dec. 11, 1925 The mixt. is compared with a graduated color tube contg coal dust at one end and stone dust at the other, with gradual intermediate mixing.

Testing mixtures of coal dust and stone dust in mines, etc. EMIL WITTE Ger. 514,162, June 15, 1929 Addn to 472,529 (preceding abstr.).

25—DYES AND TEXTILE CHEMISTRY

L. A. OLNEY

Biphenyl: A starting point in organic synthesis. GILBERT T. MORGAN AND LESLIE P. WALLS. *J Soc Chem Ind* 50, 94T(1931)—Biphenyl, now available industrially at a moderate price, is pointed out as a possible source of intermediates for synthetic dyes and drugs. Examples are given. H. W. LEAHY

The copperas indigo vat. J CLAROU *Tids* 9, 247-51(1931)—A description of the vat dyeing of cotton with indigo in the presence of FeSO_4 . A PAPINEAU-COUTURE

Troublesome fibers in the dyeing of garments. VALETTE. *Tids* 9, 141, 143 (1931)—A brief outline of the chem and tinctorial properties of some fibers which give trouble to dyes and cleaners: acetate rayon, acetylated wool, "immunized" cotton, "passive" cotton, "amido" cotton. A PAPINEAU-COUTURE

Cause of the greenish shade of cotton fabrics jigged in indigosol O by the steaming process. LOUIS A. DRIESSEN *Tids* 9, 139(1931)—One method of application of indigosol O consists in jigging cotton fabric in a bath contg indigosol O, NH_4CNS , NaClO_3 and NH_4VO_3 , drying and steaming 1-15 min. In some cases a greenish shade is produced. It is shown that this is due to the use of too much NH_4CNS which, on heating (especially if prolonged) gives rise to the formation of persulfocyanic acid. The trouble can be eliminated by reducing the proportion of NH_4CNS or by using NH_4 oxalate or lactate in place of NH_4CNS . A PAPINEAU-COUTURE

Even dyeings on wool. ALB BYTNER. *Rev chim ind*, 39, 262-5(1930).

P. THOMASSET

Partial alkaline hydrolysis of acetyl silk as a pretreatment for dyeing. W. COLTON, H. I. WATERMAN AND I. G. WOLF. *Z. anorg Chem* 44, 163-8(1931)—Superficial hydrolysis of acetylcellulose fibers in dil NaOH and AcOH , at 60° , attained max. effects, which depended on the alkali concn. The wt losses were 4.2, 6.3 and 8.4% in 0.0067, 0.0101 and 0.0133 *N* NaOH . If complete deacetylation of cellulose units took place (as has been claimed), the cellulose formed should be calculable from the wt. losses. Extn of the treated fibers with Me_2CO gave greater residues of cellulose than the calcd values. It is concluded that intermediary layers of incompletely deacetylated material exist below the completely hydrolyzed surfaces. The treated fibers were dyed under standard conditions (1 hr at 70°) with benzopurpurin 4B. Absorption was detd by colorimetric examn of the bath before and after dyeing. Absorption on the untreated fiber was negligible; it increased rapidly, at first, with increasing deacetylation, but reached a max when the wt. losses were 2-3%. Greater absorption may be produced only by lengthening the time of dyeing. Consideration of this factor is of great importance to the dyer. K. H. ENGEL

Unit heaters prevent condensation in dye houses. J. R. COOPER. *Heating, Piping and Air Conditioning* 2, 767-9(1930)—A discussion of the problem of air transportation in plants where moisture removal is of paramount importance and a description of several specific installations of unit heaters in dye rooms. E. I. S.

Note on the boiling of cotton preliminary to bleaching. J. DUTREUILLES. *Tids* 9, 245, 247(1931)—It is considered that the most logical procedure is as follows: treatment with hot 1.5% HCl , treatment with C_6H_6 or similar fat solvent and boiling at atm. pressure with Na_2CO_3 , if necessary a pancreatin treatment can be carried out between the fat extn. and soda boiling. A. PAPINEAU-COUTURE

The use of liquid chlorine in the industry of bleaching cotton cloth. LA BLANCHISSERIE ET TEINTURERIE DE THAON-LES-VOGES (Sealed communication of 1914) *Bull. soc. ind Mulhouse* 97, 60-2(1931)—A process for the prepn. of hypochlorite liquors

and their use in bleaching cotton are described. The app. consists of a column in which an ascending stream of Cl gas meets a spray of NaOH soln. or milk of lime. J. H. O.

The generating and application of x-rays for textile investigations. HERMAN BOXSPA. *Am. Dyestuff Repr* 20, 163-6(1931).—Seven references are listed.

A. K. JOHNSON

Physical and chemical properties of textiles. II. A. J. HALL. *Textile Mercury* 84, 239(1931), cf. C. A. 25, 2296.—Some effects of the phys. conditions of rayons and the methods of manuf. on the resistance of cellulose to snail enzyme degradation are shown.

A. K. JOHNSON

Identification and analysis of rayons. II. TATU. *Russa* 6, 175-87(1931).—A critical review of methods proposed for the identification and analysis of nitrocellulose rayon, the identification and analysis of acetate rayon, the sepn. of viscose from cuprammonium rayon, microscopical examn. and ultra violet examn. Bibliography of 40 references.

A. PAPINEAU-COUTURE

Pieces of artificial crepes containing acetate rayon. A. CAISSE. *Russa* 6, 243-7(1931).—Practical operating directions.

A. PAPINEAU-COUTURE

Cellulose amyloid and the utilization of the amyloid effect in the textile industry. JUSTIN-MUPPLE. *Tida* 9, 129-33(1931).—The optimum conditions for the production of amyloid consist in dissolving 5.82 parts of cellulose in 100 parts of H_2SO_4 (d. 1.580) to a thick starch like paste and then pouring the soln. slowly into a large vol. of cold H_2O , which gives a white, curdy ppt. To obtain a pulverulent ppt. it is necessary to use a large excess of acid, or to allow the latter to act for a longer period, but this results in attack of the amyloid with production of dextrin or glucose. On drying the ppt. after washing to neutrality, it becomes grayish, hard, horny and more or less translucent, but softens again on prolonged contact with cold water. The wet ppt. or its NaOH soln. does not reduce Fehling's soln., but if allowed to stand after boiling with Fehling's soln. the surface becomes very slightly reddish, showing a very slight reduction by surface action. The chief characteristic is that the dried ppt. is more resistant to the action of reagents than the undried ppt. indicating that there is probably a certain amt. of polymerization on drying. The production of cambic effects on cotton by treatment with H_2SO_4 and the parchmentizing of paper are both considered to be due to a superficial amyloidation of the fabric or paper. Rapidly passing rayon (viscose, acetate, cuprammonium) fabrics through H_2SO_4 (50-52° Bé), followed immediately by thorough washing and final treatment with dil. NH_4 , increases the tensile strength, but gives a rougher touch.

A. PAPINEAU-COUTURE

The behavior of the aqueous solution of domestic cocoons. I. II. IIMBO KANEKO. *J. Agr. Chem. Soc. Japan* 6, 623-32, 1015-22(1930).—The greater the surface tension of the aq. soln. of boiled cocoons is, the better the reeling. The iso-elec. pt. of aq. solns. in which the reeling is easy is 4.25-4.4. The sp. elec. cond. (k) of the aq. soln. = $\log C + B$, in which C is the concn. and B is the const. (about 1.1), k increases with the temp. The cond. of the soln. is changed more by acids than by bases. The cond. of a soln. having good reelability is not changed by the addn. of electrolytes.

Y. KIHARA

Medulla in wool. A new test for detection of hairiness in the fleece. B. L. ELMICK. *New Zealand J. Agr.* 42, 91-8(1931).—Medullated fibers present a bright chalky appearance when a sample of wool, free of grease and dirt, is immersed in benzene and is viewed against a black background. Pure wool fibers have a silky transparent appearance. The test can be carried out in about 5 min.

K. D. JACOB

Recovery of spinning oil from fulling plants. A. BEYER. *Tida* 9, 133-7(1931).—A brief outline of the advantages and of the various processes of recovery, and also of the necessity of controlling the compn. of the effluent to eliminate portions contr. such small amts. of oil as to render the recovery uneconomical.

A. PAPINEAU-COUTURE

Glue and gelatin in the textile industry. HERMANN STADLINGER. *Kunststoffe u. Leim* 28, 81-9(1931).—Numerous uses of these materials in various operations in mfg. textiles are described, with notes on the qualities required and suggestions for new uses. Many working formulas are included.

C. J. SCHOLLFERNBERGER

Proteins of wool (SAITO) 11A. Water requirements in the textile industry (KEHREN, STOMTEL) 14. Aromatic amines [dyes] (Ger. pat. 515,758) 10. Substitutes for sugars [in printing or dressing textiles] (Ger. pat. 518,196) 18.

Deutscher Färberkalender, 1931. Wittenberg. A. Ziemsen Verlag M. 5. Reviewed in *J. Soc. Dyers Colourists* 47, 82, *Tida* 9, 211(1931).

Dyes. CHEMISCHE FABRIK VORM SANDOZ Brit 338,858, Dec. 20, 1928 An arylaminoanthraquinone contg in the benzene ring to which is attached the arylamino group a halogen atom in β position and in the other benzene residue a sulfonic acid group is treated with sulfite (with or without the presence of phenol or a similarly acting compd) which effects replacement of the β halogen atom by a sulfo group. Examples are given of the production of water sol. dyes suitable for wool dyeing and printing giving fast blue to greenish blue shades.

Dyes. I G FARBENIND A G Brit 338,747, Dec. 4, 1928 Halogenated dibenzopyrenequinones are made by causing halogen or a halogenating agent to react in an alkali $AlCl_3$ melt with a 3,4,8,9-dibenzopyrene 5,10-quinone or a 4,5,8,9-dibenzopyrene 3,10-quinone, or with a ketone capable of being transformed into such a pyrene quinone by ring closure. The products may be after treated with a hypochlorite to improve the shades of the dyes. Several examples are given, including dyes giving orange and scarlet tints.

Dyes. I G FARBENIND A G Brit 339,489, Dec. 6, 1928 4,5,8,9-Dibenzopyrene 3,10-quinones contg both Cl and Br are made by chlorination of brominated dibenzopyrenequinones or by bromination of chlorinated dibenzopyrenequinones. Several examples are given for the production of dyes producing various red tints.

Dyes. I G FARBENIND A G Fr 37,120, June 22, 1929 Addn to 670,562 N-Dihydro-1,2,2',1'-anthraquinoneazepines contg Me groups in the anthraquinone ring are made resistant to Cl by a treatment with H_2SO_4 and MnO_2 or other oxidizing agents having a similar action. An example is given of the treatment of the 3,7'-dimethyl deriv. Cf C A 24, 2608.

Dyes. I G FARBENIND A G Fr 37,442, Sept. 18, 1929 Addn to 667,108 The sulfonic acids of Fr 667,108 (C A 24, 1223) are replaced by the corresponding carboxylic acids of aromatic diamino compds. Examples are given.

Dyes. SOC ANON POUR L'IND CHIM A BALE Fr 37,100, July 4, 1929 Addn to 677,782 (C A 24, 3378) Dyes of the general formula 2,3-(R^1N N)(HO,S) $C_{10}H_6$ (NH X NH R^1N NR^1) 6, in which X represents a residue such as CO or CS or a triazine, quinoxaline, pyridazine, pyrimidine or like heterocyclic residue, R^1 is the residue of anthranilic acid or a deriv thereof, R^1 is an aryl ring and R^1 is an aryl ring contg an OH and COOH group o to one another, are prepd by coupling the aminonaphthol derivs with a diazo compd corresponding to the group R^1 . Examples are given.

Azo dyes. I G FARBENIND A G Brit 338,672, Oct. 10, 1928 Azo dyes (examples of which dye cotton yellow and red shades) are made by diazotizing the compds produced by subjecting mono- or di-nitrobenzoylated diaminoarylsulfonic acids to slight reduction (such as is required to convert an aromatic nitro compd into an azo- or azoxy-compd, as described in Brit. 313,636, C A 24, 1227) and coupling with a coupling component. Several examples are given.

Azo dyes. I G FARBENIND A G Fr 37,157, June 28, 1929 Addn to 677,491 (C A 24, 3116) Azo dyes are prepd by coupling a hydroxydiazo compds with hydroxyquinolone derivs and choosing the reaction components in such a manner that the product contains at least one sulfonic group. Thus, 4-chloro-2-aminophenol-6-sulfonic acid is diazotized and coupled with 1-phenyl-4-hydroxy-2-(1)quinolone to give a dye which dyes Cr-mordanted wool a Bordeaux red shade. Other examples are given.

Azo dyes. I G FARBENIND A G Fr 37,203, July 10, 1929 Addn to 674,638 (C A 24, 2892) Azo dyes are prepd in substance or on the fiber from condensation products of diazo compds with primary amines which contain groups giving soly in water. Thus the diazoamino compd of 2,5-dichloro-1-aminobenzene and 4-sulfo-2-aminobenzoic acid is coupled with 2',5'-dimethoxyaniline 2-hydroxynaphthalene-3-carboxylic acid, the dye formed giving a yellowish brown shade. Several examples are given.

Azo dyes. I G FARBENIND A G Fr 37,383, Aug. 31, 1929 Addn to 645,771 (C A 23, 2041) Azo dyes are prepd by coupling 2,3-hydroxynaphthoyl 2-amino-3-naphthol ethers with diazo compds. of *p*-aminoaryl ethers, the 2 and 5 positions of which, with respect to the NH_2 group, are substituted but which do not contain COOH, SO_3H or nitric groups, e.g., 3-methyl-6-chloro-4-aminodiphenyl ether.

Metal compounds of azo dyes. SOC ANON POUR L'IND CHIM A BALE Ger 518,016, Sept. 8, 1928 See Brit. 297,478 (C A 23, 2836).

Azo dyes containing chromium. FRITZ STRAUB and HERMANN SCHNEIDER (to Soc. anon. pour l'ind chim a Bale) U S 1,798,068, March 10. A chromable dye such as that formed from diazotized 6-nitro-2-amino-1-phenol-4-sulfonic acid and β -naphthol is treated with Cr hydroxide in the presence of a water sol. salt of hydro sulfonic acid such as Na_2S . Various other examples are given. Cf C A 24, 5164.

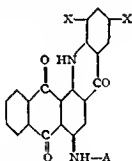
Chromium compounds of azo dyes. I G FARBENIND A-G Brit. 338,595, Aug. 22, 1929. Compsd dyeing wool greenish yellow shades are formed by boiling a Cr salt with dyes such as those formed from 2-chloroaniline (or 4-chloro-3-toluidine) and the methylpyrazolone from 2-amino-2'-hydroxy-4-sulfo-3'-carboxy-5'-methylidiphenyl sulfone or similar components Cf C A 24, 1746

Vat dyes. I G FARBENIND A-G Brit 339,396, Sept. 28, 1929. An acylating agent is caused to react in a condensation product obtained by treating a 1-(*Bz*-1-benzanthronylamino) 4 or 5- or 8-aminoanthraquinone or a substitution product halogenated in the benzanthrone residue with an alkali condensing agent. Products are obtained dyeing olive shades fast to Cl, and several examples are given. Among the acylating agents which may be used are benzoyl chloride and cinnamic acid chloride

Vat dyes. BRITISH DYESTUFFS CORP., LTD., and ANTHONY J HAILWOOD Fr 37,360, Aug 22 1929 Addn to 638,387 (C A 23, 286) The dye obtained by the fusion of naphthalimide or acenaphthenequinone with NaOH under conditions other than those indicated in the 1st example of Brit 26,690/13 is brought into a sol or colloidal form for use in the vat by treatment with an acid agent such as oleum or Cl-HSO_3 followed by a fresh pptn by diln

Vat dyes. I G FARBENIND A-G Fr 37,173, July 2, 1929 Addn. to 669,740 (C A 24, 1990) The action of Br on dibenzanthrone or isodibenzanthrone or their derivs. for the production of dyes is carried out in the presence of Cl-HSO_3 at temps below 50°

Vat dyes. I G FARBENIND A-G Fr 37,181, July 2, 1929 Addn to 674,639 (C A 24, 2891) Vat dyes are prepd by the reaction of a halogenated anthraquinone-acridone contg 3 or more halogen atoms with a halogenated aromatic amine. The products obtained probably correspond with the formula



in which X represents any halogen atoms and A is an aromatic ring. Examples are given

Vat dyes. I G FARBENIND A-G Fr 37,309, Aug 6, 1929. Addn to 679,912 (C. A. 24, 3909) Vat dyes are prepd by condensing an α hydroxyanthracene with an α -compd of a 4-methyl-5-halo-7-alkoxyisatin. Thus, α -anthrol heated in PhCl with 4-methyl-5-chloro-7-methoxyisatin gives 2-anthracene-4'-methyl-5'-chloro-7'-methoxy-2'-indole indigo which dyes fiber a greenish blue shade

Vat dyes. I G FARBENIND A-G (Wilhelm Eckert and Otto Braunsdorf, inventors) Ger 517,195, June 19, 1928 See Brit 313,887 (C. A. 24, 1296).

Vat dyes. I G FARBENIND A-G (Hugo Wulff and Ernst Honold, inventors). Ger 517,442, Apr 30, 1929 Addn to 489,957 (C A 24, 2305) Vat dyes are obtained by treating polyanthraquinonviaminobenzanthrone compds which have an anthraquinonylamino residue in the *Bz*-1-position, with alkali and, optionally, treating the resulting dye with acid condensing agents. Thus, di(1'-anthraquinonyl)-6-*Bz*-1-diaminobenzanthrone (obtained from 6-*Bz*-1-dihalo-benzanthrone and 1-aminoanthraquinone) is fused with NaOH and alc at $180-200^\circ$. The resulting mass is heated with water, filtered and washed. The dye so obtained colors vegetable fibers in olive-brown shades from a NaOH and hydrosulfite vat. Further examples are given Cf. C. A. 24, 4941

Vat dyes. I. G FARBENIND A-G. (Georg Rösch, inventor) Ger. 518,017, Aug 9, 1929 Yellow to orange vat dyes are prepd by combining 2 mols of an α -aminoanthraquinone with 1 mol of naphthalene-2,6-dicarboxylic acid or its derivs not contg OH groups. Suitably, the dichloride of the acid is heated with the α -aminoanthraquinone in PhNO_2 . Examples are given

Vat dyes. I G FARBERND A G (Heinrich Nettesheimer and Wilhelm Schneider inventors) Ger 518,230, July 26, 1929 The dyes described in Ger 491,430 (C A 24, 248) are halogenated. The products are faster to Cl than the initial dyes and give modified dyeings. Examples are given.

Violet vat dyes. I G FARBERND A-G (Maximilian P Schmidt and Wilhelm Neugebauer inventors) Ger 502,352 Apr 30, 1927 Haloperylenetetracarboxylic acids contg not more than 3 halogen atoms are condensed with *o*-diamines. An example is given.

Blue-green vat dyes. SOC. ANON FOUR ETVD CHIM A BALE Brit. 339,516, Dec. 24 1929 2 Methyllenzanthrone or a deriv contg a replaceable substituent in the Br-1 position is sulfonized and converted into products of improved properties by dissolving in H_2SO_4 and dilg the soln thus formed continuously with water or with aq H_2SO_4 so that the dye is pptd but impurities are retained in soln, or purification may be effected by reducing to the keto compd, sepg, and oxidizing the latter if desired.

Blue vat indanthrone dye. D A W FAIRWEATHER, J THOMAS and SCOTTISH DYES LTD Brit 338,891 May 23 1929 2 Aminoanthrahydroquinone 9,10-disulfonic ester is oxidized by use of a cupric salt such as the sulfate or chloride (suitably with addn of NaOH or HCl) to avoid danger of over-oxidation and produce a product which is presumably indanthrone.

Acid wool dyes. I G FARBERND A-G (Rudolf Muller and Ernst Honold, inventors) Ger 518,229, Mar 19, 1928 The diazo compds. from 1 aminoanthraquinone-2 sulfonic acids substituted in the 4 position with a readily exchangeable atom or group or the residue of a base are treated with primary or secondary bases. Thus, the diazo compd from 1 amino-4 bromoanthraquinone-2-sulfonic acid may be treated with $PhNH_2$ while cooling in the presence of a little Cu. The reaction proceeds with evolution of N. Other examples are given also. Cf C A 25, 214.

Acid wool dyes. WILHELM ECKERT (to General Aniline Works) U S 1,790,011, March 10 Dyes producing generally yellow or greenish yellow dyeings (according to the several examples described) on animal fibers or cellulose esters and ethers are produced by causing a compd of the group consisting of NH_2 , alkyl- NH_2 , aralkyl- NH_2 , aryl- NH_2 , and *o*-arylene (NH_2)₂ (the aryl residues being substituted or not) to act upon a compd of the general formula of a naphthalic anhydride wherein may be substituted the sulfonic or the sulfamic acid group and a univalent substituent in the presence of water at a temp of about 100°. The use of the Na deriv of 4-sulfamino-1,8-naphthalic anhydride and o -C₆H₄(NH₂)₂, and similar starting materials is described, and the same or similar dyes are obtained by transforming the reaction products of naphthalic anhydride or a deriv thereof, with NH_2 , an aliphatic or aromatic amine into the sulfonic acids or sulfamic acids, resp. or by subjecting naphthalic anhydride or a deriv thereof to react with a sulfonic acid of an amine. U S 1,790,012 describes the production of generally similar dyes by treating nitronaphthalic acid derivs. (such as nitro-1,8-naphthalimide contg H, alkyl, aralkyl or aryl on the N) in an aq soln with such reducing agents as simultaneously effect the entrance of a sulfonic acid group. Such agents are, for instance, hyposulfite, bisulfite or the like. It is advantageous to add to the aq soln a water sol. org. solvent, *e. g.*, alc. The reaction is generally carried out at boiling temp, but it is also sufficient to heat the soln. to a somewhat lower temp. It is also possible to use as parent materials such nitronaphthalic acid derivs. as contain a sulfonic acid group. In this case the introduction of a further sulfonic acid group is unnecessary and it is sufficient to reduce the nitro group. Examples are given of the use as starting materials of *N*-methyl-4-nitronaphthalimide (or the corresponding ethylamide or phenylamide), 4-nitronaphthalimide and *N*-(*p*-sulfophenyl)-4-nitro-1,8-naphthalimide. Cf C A 25, 1097.

Anthraquinone dyes. I G FARBERND A-G Brit. 339,266, Sept. 3, 1929 1-Amino-4 bromoanthraquinone-2 sulfonic acid is caused to react with an aromatic amine contg a hydroxyl group, which is etherified by a polyhydric alc., such as *o*- or *p*-aminophenol glycol ether or *p*-aminophenol glycerol ether, suitably in the presence of Cu or a Cu salt and of Na₂CO₃. The dyes dye wool clear blue or greenish blue shades.

Acylaminoanthraquinone dyes. I G FARBERND A-G Brit. 339,267, Sept. 3, 1929 Two mol proportions of *o*-aminoanthraquinone or a substitution product are condensed with 1 mol proportion of a reactive deriv of a naphthalenedicarboxylic acid, a diphenylmethanedicarboxylic acid, diphenyl *p,p'*-dicarboxylic acid or a halogen substitution product of such compds (with the exception of such dicarboxylic acids as are capable of forming internal anhydrides). Various examples are given for the production of dyes giving yellow, orange and red shades.

Dibenzanthrone dyes. R. S BARNES, R. F THOMSON, J THOMAS and SCOTTISH

DYES, LTD. Brit. 339,321, May 31, 1929. Fluoride esters of dibenzanthrone or halogenated dibenzanthrone are treated with oxidizing agents (such as $\text{Na}_2\text{Cr}_2\text{O}_7$, $\text{K}_2\text{S}_2\text{O}_8$, KMnO_4 , V_2O_5 or H_2O_2) in the presence of an inorg. acid of at least 4% strength. Various details and examples are given.

Sulfurized dyes. 1. G. FARBERND, A.-G. Fr. 37,484, Oct. 2, 1929. Addn. to 633,783 (C. A. 23, 3810). Sulfurized dyes are prepd. by heating halogenated dinaphthylene dioxide with S, with or without diluents of high b. p., or melting these compds. with polysulfides in the known manner. Examples are given of dyes from the tetrachloro, pentachloro, tetrabromo and hexachloro compds. Cf. C. A. 25, 1090.

Sulfuretted dyes. 1. G. FARBERND, A.-G. Brit. 330,410, Oct. 10, 1928. Acid mordant dyes probably of the thiazine series are obtained by reaction of S in fuming H_2SO_4 on a β -hydroxynaphthoquinone-sulfoarylimide (in which O, OH and the NRSO_2H group are in the 1, 2 and 4 positions, resp.) in which the aryl group (R) may contain further substituents with the exception of alkylamino groups. Examples are given for the production of dyes dyeing wool from an acid bath reddish violet shades which turn green when after-chromed; the dyes are suitable for printing wool or cotton by the chrome printing process. Cf. C. A. 25, 2005.

Pastes from insoluble mordant dyes. E. YA VINETSKAYA. Russ. appl. 22,888, Sept. 13, 1927. Pastes are prepd. from alizarin by mixing the dry dye with 5–10% of a protective colloid, such as gum tragacanth, etc., and grinding the mixt. in a colloid mill together with water.

Pasting sulfur dyes with resin soaps. 1. G. FARBERND, A.-G. (Franz Schweitzer-Hennig and Walter Häge, inventors). Ger. 518,198, June 10, 1928.

Improving the solubility of dyes. ARTHUR R. MURPHY and JOSEPH B. ORSEN (to Newport Chemical Corp.) U. S. 1,793,113, March 10, 1929. A dye such as direct fast yellow 8G is dissolved in water in the presence of a sol. pyrophosphate salt such as $\text{Na}_2\text{pyrophosphate}$, which serves as a solubilizing agent.

Fuchsin. C. N. IGARYEV and I. I. VASIN. Russ. appl. 50,938, July 11, 1929. Fuchsin is prepd. by heating above 100° in the presence of air a mixt. of aniline, *o* and *p*-toluidine and activated carbon without the addn. of any kind of oxidizer. The fuchsin is extd. from the melt by any of the known methods.

Parafuchsin. S. N. IGARYEV and I. I. VASIN. Russ. appl. 50,307, Nov. 29, 1929. Parafuchsin is prepd. by heating a mixt. of (β - $\text{H}_2\text{NC}_6\text{H}_4$) $_2\text{Cl}_2$, aniline and its salt, iron chloride and activated wood charcoal in the presence of air. The dye formed is extd. from the melt, which is first freed from the excess of aniline by the usual methods. The extn. can be carried out, e. g., first with acidified water and then with org. solvents.

Bisulfite compound of alizarin blue. A. A. RABINOVICH. Russ. appl. 49,402, June 19, 1929. A mixt. of alizarin blue paste or the dry dye with a soln. of alkali is super-satd. with sulfurous gas under agitation and let stand under ordinary or superatmospheric pressure at ordinary or elevated temp. until a sample is completely sol. in water. The product is then isolated by ordinary methods.

Anthranthrene derivatives. 1. G. FARBERND, A.-G. (Max A. Kunz, Karl Köberle and Erich Berthold, inventors). Ger. 518,234, Aug. 10, 1928. New compds. useful as *rot dyes* or *intermediates* are prepd. from di- or poly-aminoanthranthrones by substituting a radicle for H in one or more of the NH_2 groups. Examples are given describing the reaction of diaminoanthranthrone with PhCOCl , 1-chloroanthraquinone and 1-aminoanthraquinone-2-carboxylic acid chloride. Cf. C. A. 25, 602.

Condensation products from benzoïn. SOC. ANON. ROY. LTD. CHIM. A. RILEY. Ger. 517,498, July 20, 1928. Fr. 37,391, Sept. 5, 1929. New compds. are prepd. by treating benzoïn with aromatic sulfonic or carboxylic acids in the presence of a condensing agent, e. g., H_2SO_4 , oleum, or $\text{H}_2\text{SO}_4\text{Cl}$. Thus, a soln. of Na benzenesulfonate in H_2SO_4 may be stirred with benzoïn at 80 – 90° until a sample gives a clear soln. in water. Alternatively, mixts. of benzoïn and aromatic compds. may be treated with reagents having both a sulfonating and condensing action. Thus, $\text{C}_6\text{H}_5\text{I}$ and benzoïn may be fused together and the cooled melt introduced into H_2SO_4 and $\text{H}_2\text{SO}_4\text{Cl}$ at 60° . Other examples are given also. The products are useful as assistants in the *rouge* of lakes and dye pastes and in dyeing with ice colors and vat dyes.

Dye intermediates. 1. G. FARBERND, A.-G. Brit. 338,764, Dec. 10, 1929. 2. Arylbisanthrones are made by treating benzanthrones contg. the group $-\text{CH}(\text{R})\text{CN}$ (in which R represents aryl) in the 2 position, described in Brit. 319,503 (C. A. 24, 2613), with oxidizing agents such as KMnO_4 or $\text{Na}_2\text{Cr}_2\text{O}_7$ in the presence of org. diluents such as glacial HIO_4 or acetone. Several examples are given.

Dye intermediates. 1. G. FARBERND, A.-G. (Walter Mieg, Berthold Stein and

Willy Trautner, inventors) Ger 518,213, May 3, 1927. New condensation products are prepd by treating aminoanthraquinones or their derivs or substitution products with alc. alkali hydroxide solns at a temp below 100° while leading air into the reaction mixt. The products are strongly colored and are sol in H_2SO_4 but they are not useful as vat dyes. 1 example is given.

Gray dyeings on furs, hairs and leathers. I G FARBEIND A-G. Brit 339,444, Nov 7 1928. The material (with or without mordanting) is treated with a soln contg a suitable oxidizing agent such as H_2O_2 and a 4-amino-N-alkyl or N-alkyldiphenylamine or a 4'-halo- or 4'-alkyl deriv. Various examples and details of procedure are given.

Color lakes. WILHELM EBELTIN. Ger 515,981, Jan 25, 1928. Color lakes are produced by pptg basic dyes by sulfite-cellulose lye. Fixing agents, emulsions, fats and waxes may be added. Thus methyl violet is treated with the lye $Al(OH)_3$ may be added to form a printing paste. Addn of a colloid (e.g., bentonite) and wax or oil renders the lake suitable for use as a marbled paper-coloring agent.

Dyeing textile materials. PIETER MIJES (to Two-Tone Corp.) U. S. 1,796,039, March 10. Various details of app and procedure are described for coloring webs of permeable fabric by forming 2 color clouds while sepg the heavy particles or drops from the lighter particles and exposing one face of the fabric to one of the clouds and the other face to the other cloud. Cf C A 24, 3393-4.

Dyeing piece goods and yarns with vat dyes. I G FARBEIND A-G. (Gerhard Poetzsch and Hermann Jager, inventors) Ger 518,107, Feb 3, 1928. Better penetration is obtained by slop padding the materials with the unvattd dye paste with the addn of an agent promoting running or bleeding, and then treating the materials as usual with an alk. hyposulfite soln. Suitable agents for promoting running are emulsions of oils and fats or their free acids or water sol derivs, e.g., Turkey red oil or cellulose derivs forming neutral or alk. aq. colloidal solns, e.g., an alkali salt of carboxymethyl cellulose. The method is intended particularly for tightly beaten cotton or linen fabrics or strongly twisted or mercerized yarns. Examples are given.

Dyeing tulle goods. V. V. KUVCHENSKII and N. V. VOZOKOV. Russ. appl. 41,382, Feb 22, 1929. Tulle goods are steamed, known dye solns are added by the usual methods and the goods are again steamed.

Dyeing acetyl cellulose. WOLFGANG JABEK and JOSEPH LANG (to Soc. anon. pour l'ind. chim. à Bâle) U. S. 1,796,023, March 10. Acetylcellulose is treated with a colloidal prepn from an intermixture of insol. aminoazo dyes having an affinity for acetylcellulose, with waste sulfite liquor, which serves as a dispersing agent. U. S. 1,796,029 relates to the similar use of aminoanthraquinones with waste sulfite liquor.

Dyeing cellulose esters. BATTISTO CELANESE, LTD. Brit. 339,429, Oct. 26, 1928. In dyeing fabrics, such as those comprising cellulose acetate, formate, propionate or butyrate on machines of the wench type, formation of creases is avoided by partially saponifying the ester before dyeing, and delustering may be simultaneously effected. Various details of procedure and reagents used are given.

Dyeing fur, hair and leathers. I G FARBEIND A-G (Karl Marx and Karl Bittner, inventors) Ger 518,199, Dec 11, 1928. The mordanted or unmordanted materials are treated with an oxidizing agent and a base (or its salt) of the formula $R_1R_2NC_6H_4NR_3CH_2CH_2N(alkyl)_2$, where R_1 is H or an alkyl group and R_2 and R_3 are H or alkyl or aryl groups. Nuclear substitution products of such bases may also be used. Thus, Cu- or Fe mordanted fur may be treated with an aq. bath contg H_2O_2 and 4-amino-1 (ω -diethylaminoethylamino)benzene, brown shades being obtained. Other examples are given also. Cf C A 25, 2007.

Printing with vat dyes. I G FARBEIND A-G. Fr 37,188, July 3, 1929. See Brit. 314,904 (C A 24, 1522).

Printing with vat dyes. I G FARBEIND A-G. Fr. 37,282, July 4, 1929. Addn to 642,991 (C A 23, 1514). Merchandise is printed with a vat dye mixed with thickening agents other than cellulose esters or ethers capable of easy pptn., and the merchandise is treated with a reducing agent and an alkali. Examples are given of the use of carrageen moss.

Printing fabrics. I G FARBEIND A-G (Ludwig Lochner, inventor) Ger 516,088, Sept 12, 1928. Addn to Ger 513,526 (C A 25, 2003). In producing polychromatic effects by ester salts of vat dyes by the method of 513,526, β -naphthol and other coupling compds are used instead of 2,3-dihydroxynaphthoic acid arylides. Thus, fabric is soaked in a lye contg β -naphthol, Turkey red oil and NaOH. For red printing, the dried fabric is treated with a paste contg diazotized β -nitroaniline, water, starch and $AcONa$. Further examples are given.

Printing wool. I. G. FARBENIND A.-G. Fr. 37,474, Sept 30, 1929 Addn to 662,205 (C. A. 24, 510) Illumination effects are obtained by printing wool, colored with a dye capable of being discharged, with discharge agents containing esterified salts of leuco compounds of vat dyes, then treating the merchandise after steaming with a solution of a persulfate containing a free acid Cf. C. A. 25, 1102

Piece-dyeing machine. SIEMENS SCHUCKERTWERKE A.-G. (Georg Schlenk, inventor) Ger. 515,926, June 10, 1929 Addn to 459,546

Apparatus for washing, dyeing or other treatments of articles with liquids. KENNETH C. ANDERSON U. S. 1,790,312, March 17 Structural features of an apparatus suitable for treating fabrics or other materials

Treating textiles. CAMILLE DREYFUS Fr. 37,382, Aug 31, 1929 Addn to 679,429 (C. A. 24, 3911) A limited amount of a saponing agent in other than paste form is applied to textile materials of or containing cellulose esters and rapidly dried

Cleaning textiles, etc. LOUIS A. H. LANTELME Fr. 695,131, May 6, 1930 A colloidal product such as starch, paste, bone glue or skin glue is used in the cleaning, washing, bleaching, dyeing or charging of textiles or threads of all kinds

Waterproofing textile materials. FRITZ FUCHS Austrian 121,364, Sept 15, 1930. Addn to 100,167 (C. A. 22, 4836) The materials are impregnated with a solution of degradation products of vulcanized rubber in a high boiling solvent, e. g., pitch, asphalt or bitumen, with or without a drying oil The preparation of the solution may be interrupted before solution is complete the undissolved portion acting as a filler

Treating textile materials for mercerization EMIL GRIMMER U. S. 1,796,598, March 17 Woven material is subjected to the action of a shrinking liquid such as mercerizing solution and is afterward subjected to a series of alternate cooling and transverse stretching treatments

Bleaching and mercerizing machines. ALLGEMEINE ELEKTIZITÄTS GES Ger. 515,613, Aug 14, 1928 Details of regulating

Bleaching bast fibers. I. G. FARBENIND A.-G. Brit. 339,550, March 16, 1929 Bast fibers, either as crude fiber or in yarn or finished goods, are bleached by successive use of baths comprising (1) peroxide, (2) acid solution, (3) Na_2CO_3 solution, (4) alkali or neutral solution and (5) peroxide Various details are described

Bucking process for vegetable fibrous materials. GUSTAV ULLMANN Austrian 121,528, Oct. 15, 1930 See U. S. 1,787,880 (C. A. 25, 823)

Bucking linen, etc. WERNER ERB Ger. 515,675, May 10, 1927 Vegetable fibers are bucked in a mechanically stirred bleaching bath at 60-70°

Preparing vegetable textile fibers for spinning. JOSEPH VIALLET (to N. T. Artificial Wool Co., Ltd.) U. S. 1,795,414, March 10 Cleaned waste vegetable fibrous material such as waste jute fabric is treated with a solution which promotes fermentation such as a dilute solution of MgSO_4 , KNO_3 , and K phosphate and heated to 30° with a softening agent such as olein, and after the fermentation has sufficiently proceeded the material is unravelled and boiled with a dilute alkali solution

Textile fiber from materials such as ramie, flax, hemp, sisal and bamboo. DIN-SHAW R. NANJ (one-half to Duncan M. Stewart) U. S. 1,796,718, March 17 The material is heated with a 0.1-0.5% $(\text{NH}_4)_2\text{SO}_4$ solution at a temperature corresponding to a vapor pressure of 1-2 atm for 60-20 min and then subjected to pressure to remove slimy matter; subsequently the fiber is washed, dried and obtained from the straw by an ordinary scutching or similar process U. S. 1,796,719 describes a process for the manufacture of textile yarns consisting in resolving natural bast fibers such as those of plants of the genus *Boehmeria* embodying ultimate filaments of a length of at least 2.5 in. into such ultimate fibers, blending the latter with synthetically produced fibers such as any variety of rayon and then spinning the blend Cf. C. A. 24, 1750

Fiber and yarn from unretted flax. HENRY C. WATSON and MARTIN WADDELL (to Watson-Waddell Ltd.) U. S. 1,795,528, March 10 After breaking, scutching and twisting the fibers into a rove, the rove is wound onto an open reel in the roving frame, retted while on the reel, washed and spun U. S. 1,795,529 and U. S. 1,795,530 also relate to apparatus for treating flax

Dressing yarns. EDOUARD PINEL Ger. 518,195, Nov. 27, 1926 Yarns of cotton, wool, etc., are treated with a solution of carob bean gum, which may contain a softening agent such as glycerol, and then with a solution of starch or other size

Singeing fabrics. W. OSTHOFF Brit. 339,281, Aug 28, 1929 A rich excess of O is supplied with the primary air or O to a gas or vapor burner and singeing is effected by the highly heated O thus obtained. Apparatus is described. Cf. C. A. 24, 511

Delustering fibers of fabrics of regenerated cellulose. I. G. FARBENIND A.-G. (Walther Schieber, Hugo Pfannenstiel and Herbert Mahn inventors) Ger. 518,194,

Aug. 22, 1929 The materials are treated with a soln. of a cellulose ester or ether in conjunction with a non solvent for the ester or ether, with the result that a cloudy film is produced on the materials when they are dried. Thus, regenerated cellulose yarn may be treated with a 0.2% soln. of benzylethylcellulose in equal parts of MeOH and $C_{12}H_{26}$ to which 10% of water has been added. Alternatively, the yarn may be impregnated with a simple soln. of the cellulose deriv., and the non solvent applied later *et seq.* by steaming.

Carbonizing cellulosic fibers in mixed goods. H. DREYFUS. Brit. 339,300, Aug. 2, 1929. Mixed fabrics contg. cellulosic fibers such as cotton together with other materials such as cellulose derivs., wool and silk are heated in inorg. acid or acid salt solns. sufficiently dil. to have no deleterious effect on the non-cellulosic materials, to render the cellulose fibers removable without drying in the presence of a carbonizing reagent. Ornamental effects may be produced by use of colored or other reserves, and delustering of cellulose acetate may be prevented by avoiding undue heating or by using suitable protective agents.

Filling and weighting fabrics of cellulosic materials. R. W. R. MACKENZIE, E. H. ROBINSON, D. LUNDEN and M. FOAR. Brit. 339,651, Sept. 23, 1929. Materials such as cotton or linen are loaded and filled "to an invisible way" with mineral loading agents such as clay, talc or barite (alone or together with unwoven cotton, wood fiber, paper or paper pulp) by rendering them sticky by treatment with a plasticizing agent such as H_2SO_4 , caustic alkali, CS_2 , cuprammonium soln. or the like, applying the filling before, during or after the plasticizing treatment, and subsequently subjecting the material to a setting bath. The properties of the material are stated to be improved in various respects.

Artificial silk, etc. I. G. FARNHIND A.-G. Brit. 339,089, Nov. 14, 1928. Artificial threads, hairs, bands, etc., are made by dissolving natural silk waste in H_3PO_4 , coagulating the viscous soln. thus formed by use of an alkali metal salt soln. (suitably a NaCl soln. contg. also NH_4 formate) and stretching the coagulated material to the desired form. Various details of the process are described.

Artificial silk, etc. H. DREYFUS. Brit. 338,983, May 31, 1929. Filaments, threads, ribbons, films, etc., of cellulose or its derivs. such as nitrocellulose are subjected to the action of sulfonic acids contg. at least 6C atoms (such as benzenesulfonic acid) of a concn. of at least 20% (preferably 30-50%) for the purpose of improving the strength and elasticity and modifying the luster of the material. The process is particularly suitable for the treatment of viscose products made from alkali cellulose unripened or ripened for only a short time. After treatment, the sulfonic acids are preferably removed by washing with water alone or with water contg. an alkali or base. As a substitute for or addn. to the sulfonic acids, acids such as methylphosphonic acid or nitrophenylphosphonic acid may be used, and benzyisulfonic acid, phenolisulfonic and disulfonic acids and naphthalene- and naphtholsulfonic acids may be used. Various details and modifications of procedure are described.

Dull finish viscose silk. BENNO BORZYKOWSKI (to Swiss Borvisik Co.), U. S. 1,796,744, March 17. The viscose silk, while still contg. S, is treated directly with a bleaching agent such as a very dil. Cl soln. which will not remove the S.

Treating raw silk. I. G. FARNHIND A.-G. (Robert Gressbach and Otto Ambros, inventors). Ger. 513,373, Jan. 22, 1928. Raw silk is scoured by treatment with a soln. contg. an activated protease, *e. g.*, papain activated with a trace of HCN. Examples are given.

Impregnating jute bags or other fabrics. L. D'ANTAL. Brit. 338,538, Oct. 19, 1928. An impregnating material is prep'd by oxidizing a vegetable drying oil (such as linseed oil or sunflower oil) or a mixt. of such oils by treating at temps. of 180-200° or higher with acid sludge from oil refining. Various details of treatment are described and mention also is made of the use of castor oil and substances contg. the SO_4 radical.

Apparatus for carbonizing woollen fabrics in open width. ERNST GESSNER A.-G. Ger. 518,310, May 6, 1928.

Gas-cell fabric for aircraft lighter than air. ARTHUR D. CUMMINGS (to Goodyear Tire & Rubber Co.) U. S. 1,797,189, March 17. A textile material is impregnated with waterproofing material such as Al acetate, oleate or stearate and provided with a coating of an agglutinant such as glue or gelatin and with a ply of material, such as goldbeater's skin or "cellophane," which is highly impermeable to inflating gases, over the agglutinant.

Nitrocellulose and linocryn mixture suitable for coating fabrics, etc. VICTOR SCHOLZ (one-half to Atlas Agro-Chemische Fabrik A. G.) U. S. 1,796,219, March 10. Linseed oil is oxidized and polymerized in the presence of a Pb comp'd. catalyst such as

PbO or Pb linoleate and kneaded, the kneaded oil is dissolved in air, with use of heat and pressure, and the resulting soln is mixed with nitrocellulose. *Cl. C. A.* 23, 2713; *Mothproofing fabrics.* *Tun LARVEX COKE.* *Ger.* 615,612, Nov. 11, 1928. See *Brit.* 301,121 (*C. A.* 23, 4085).

Mothproofing waxes. I. G. PARNPOND A.-G. (Wilhelm Lommel and Heinrich Mäurer, inventors). *Ger.* 615,950, Jan. 1, 1930. The waxes are rendered mothproof by treatment with Sr salts, e. g., $\text{Sr}(\text{NO}_3)_2$.

26—PAINTS, VARNISHES AND RESINS

A. H. RABIN

Aluminum powder for paints. *Brit. Eng. Standards Assoc., Specification No. 388,* 6 pp (Aug., 1930). Specifications. R. I. S.

Measurement of color tone. W. VAN WOLLEN SCHOLTEN. *Farben-Ztg.* 35, 1821-6 (1930).—After a brief discussion of the Ostwald and the Baumann-Prese systems of color designation, the author recommends the step photometer. It should be useful in standardizing tinting pigments. G. G. SWARD

Luminescence analysis. II. Luminescence of white painters' colors and the employment of luminescence analysis for investigations of paintings. ERNST REUTER, AND ARTHUR KUTZLEHN. *Monatsh.* 57, 9 (1931) *cl. C. A.* 24, 2953. The work was undertaken from the point of investigation of paintings to det. their origin. Most ZnO whites showed canary yellow, one only showing brown. This was so strong in even min. quantities that quant. tests were impossible. PbCO_3 showed a beige color and could be identified in mixts. with BaSO_4 . TiO_2 showed dark violet and mixts. with ZnO showed gradations up to the ZnO yellow. In con. colors the vehicle had little influence on the luminescence. The presence of ZnO could be established with certainty, but PbCO_3 not so surely, partially on account of the variety of reactions possible after exposure to air for long periods. It varies greatly as the paintings yellow with age. GEORGE M. IVANS

Difficulties in paint and varnish manufacture. ROBERT HORTSMANN, *Farbe u. Lack* 1931, 135.—Many difficulties attributed to faulty material are often due to improper formulation or to carelessness. A careful study by producer and user will eliminate misunderstandings and uncover economies. Some specific examples are given. G. G. SWARD

Preparation and investigation of titanium white. PAUL ARKHADEV AND KURT HILTZ, *Z. anorg. allgem. Chem.* 196, 257-341 (1931).—A completely Fe-free metatitanic acid can be obtained from a Ti sulfate soln. contg. Fe by sepn. as TiK oxalate followed by hydrolysis of the mixed oxalate. This takes place at 135° , the yield depending upon the addn. Orthotitanic acid can be sepn. quantitatively in the cold at a pH of 4. At this acid concn. 1.5% remains in soln. By resolu. of orthotitanic acid in H_2SO_4 and hydrolytic sepn. of metatitanic acid at 95° , a completely Fe-free high grade paint pigment is obtained. The brilliancy of this Fe-free TiO_2 is comparable with that of other white pigments, while its covering power is not approached by any other. The influence of the ratio $\text{TiO}_2/\text{BaSO}_4$ upon covering power was studied. Pigments which were obtained by pptn. of the metatitanic acid upon a BaSO_4 suspension were superior in covering power to the corresponding mech. mixt. of TiO_2 and BaSO_4 . The best covering power was obtained when the BaSO_4 and metatitanic acid were simultaneously pptd. In the extn. of Ti from limonite with H_2SO_4 , 8 hrs.' treatment with a 65% acid is sufficient. This produces a basic sulfate metatitanic acid mixt. contg. Fe. Removal of Fe by treatment with acids leads to too great a loss of Ti. H. SPURKATZ

Standoil. R. PAIRARU. *Vestnikh* 3, No. 10, 122-8 (1930); *Chimie & Industrie* 25, 418 (1931).—Prolonged heating of linseed oil (and others) at $250-300^\circ$ produces an increase in the viscosity and the acidity, which occurs more rapidly the higher the temp. The effect of temp. was studied on 10 kg. samples in Al kettles at 250° , 285° , 290° , 295° and 300° . The viscosity increases as an exponential function of the time of heating. The acidity increases fairly rapidly at first (1st hr.), and then approx. linearly. It is preferable not to heat above 300° , as the increase in acidity is too rapid and becomes too great in the very highly thickened oils. A. PAPINPAU-CHUTIER

Colophony in varnishes and oil paints. R. SCHWARTZ. *Farbe u. Lack* 1931, 123-4.—The low price and many uses of colophony make it a very important raw material. The principles of its uses in gloss oils, rubbing varnishes, tung oil spar varnishes and other gum are briefly outlined. G. G. SWARD

Kauri gum in nitrocellulose lacquers. 1. Solubility of kauri gum in mixed lacquer solvents. EDWARD S. BRACKMAN. *New Zealand J. Sci. Tech.* 12, 65-63 (1920).—Kauri is sol in (1) anhyd to which may be added AmOAc or (and) $\text{C}_2\text{H}_5\text{Cl}$; (2) petroleum distillate used in mixts with other solvents; (3) Cellosolve. AmOH is also used with these. All these are compatible with pyroxylin. This long paper is a treatise on using kauri with these in pyroxylin lacquers. D. THURMAN.

Printing-ink patents. JOSEPH ROSSMAN. *Am. Ink Maker* 8, No. 8, 21-3; No. 9, 21-4, 41 43, No. 10, 27, 29, 41. No. 11, 27, 29, 41, No. 12, 27, 29 (1920); 9, No. 1, 25, 27 (1921). E. J. C.

TiO_2 (SPRITZEL) 18. Separation and size distribution of microscopic particles (ROLLER) 2. Use of vulcanized rubber waste as a fixing agent in bronze printing (RICE) 30. Rubber conversion product [for use in varnishes or lacquers] (U. S. pat. 1,797,158) 30. Drying apparatus for pigments (Ger. pat. 518,370) 1. Purifying pitch, etc. (Brit. pat. 339,470) 22.

JACOB J. WERNER. Die Fabrikation von Wachstuch, amerikinischem Ledertuch, Wachstafel, Maler und Zeichenleinwand, Teertuch, unverwundlichen Geweben, Linoleum, Linoleumersatz und Kunstleder. Leipzig: A. Hartleben's Verlag. 182 pp. M 4.50, bound M 5.50.

Green paint. P. N. SHUMAGIN. Russ. appl. 72,760, June 4, 1920. Dehydrated CuSO_4 is added with agitation, to a mass of white lead and litharge ground in linseed oil after heating and recoloring. The mass is then passed through a color mill, and water is added. The product is taken up in ground white lead and a soln. of rosin in gasoline.

Resist-resistant paint. CARL NITTINGER (to Carnit-Rosschutz & Farbstoff G. m. b. H.). U. S. 1,795,962, March 10. A paint suitable for use on iron, wood or other surfaces is formed of a vehicle such as varnish and a pigment comprising finely ground coal ash about 80 and finely ground "electro-positive Zn" about 20%.

Asbestos- and antifouling paints. N. J. SHERRICK. Brit. 339,128, Dec. 17, 1920. Paints are formed contg. over 25% by wt. of Hg oxide, with vehicles such as MeOH , shellac, turpentine, linseed oil and japan and pigments such as ZnO , Indian red or other Fe oxide pigments or red lead or other Pb compds.

Masking paste for use in sign painting, etc. ALBERT E. VAUGHN. U. S. 1,795,453, March 10. A pigment such as yellow ochre is mixed with an emulsion formed of honey 10, evapor. milk 5 and glycerol 6 parts and a small proportion of "sulfurated potassium."

Oil colors for printing, etc. SIGMA HIRSHMAN FENKEL. Ger. 518,322, Sept. 11, 1920. A strongly blown linseed oil having an acetyl no. of at least 100 is used as the vehicle.

Iron oxide pigments. I. G. FARRINGTON & G. (Bernhard Witzschmitt and Annemarie Reuther inventors). Ger. 515,563, Dec. 6, 1928. A $\text{Fe}(\text{OH})_3$ -contg. ppt. is oxidized by O gas in presence of water or aq. salt solns. at temps. above 100° and with raised pressure. Thus, $\text{Fe}(\text{OH})_3$ is pptd. from a FeSO_4 soln. by NaOH and the ppt. oxidized in an autoclave at 180 - 230° and 10 atm. to give a bright red pigment. Further examples are given.

Lead pigments. SWISS INVENTIONS SYNDICATE, LTD., and A. V. BLOW. Brit. 319,357, Sept. 7, 1929. A pigment consisting of a fine dispersion of Pb in Pb oxide is prepd. by melting an alloy rich in Pb (suitably an alloy contg. Pb 85, Sb 17 and Sn 2%) in a thin layer in an air tight furnace in the presence of Ca resinate (suitably 0.5% the quantity of the alloy), maintaining the temp. just above the m. p. of the alloy, and creating a vacuum in the furnace so that streams of air are drawn into the furnace to oxidize the Pb of the alloy. A temp. of 250 - 310° is suitable with the specified alloy.

Ultramarine. V. I. OSOVLITSKI and M. P. REUT. Russ. appl. 59,974, Dec. 5, 1929. The black caustic obtained from spent sulfite liquor is evapor. to dryness, disintegrated, mixed with sulfur, kaolin and silica, calcined in crucibles or muffle furnaces and then worked over by usual methods.

Composition for colored pencils. V. M. ROBINOV and I. M. FEDOROV. Russ. appl. 47,402, May 17, 1929. Graphite, talcum, etc., are added to the mass obtained in the prepn. of methyl violet by oxidation of PhMMe . The mass is then worked up by the usual methods after the addn. of a gum, such as gum tragacanth.

Drying oil products. J. SCHWENK. Brit. 338,932, May 21, 1929. Products having properties generally similar to those of wood oil are obtained by heating oils of the linseed, soy bean and fish oil classes to 200° or somewhat higher with small quantities of catalysts such as H_2SO_4 , SnCl_4 , Al_2O_3 , HgO , I bromide, Fe, Ni, Zn or Sn. The corre-

sponding fatty acids may be similarly treated and the products thus formed then esterified with glycerol or glycerol or a mixt. of these.

Driers for paints, varnishes, etc. I. G. FARMINSKY A-G. (Franz Pohl, inventor) Ger. 518,001, Oct. 20, 1927. See Brit. 311,710 (C. A. 24, 1070).

Coloring nitrocellulose lacquers and plastics. I. G. FARMINSKY A-G. Ger. 518,018, Sept. 21, 1927. Aukh to 515,037 (C. A. 25, 1092) and 510,315 (C. A. 25, 2011). The dyes specified in Ger. 515,037 and 510,315 are used in conjunction with dimethoxy phenylamines. Green colors fast to light are so obtained.

Protective coating on iron, wood or other materials. WALTER BAUER U. S. 1,705,703, March 10. Protection from the atm. is effected by applying at ordinary temp., a mixt. of castor oil and 5-10% metal oxides such as those of Pb, Mn, Co or Zn with an alkali of solvents consisting of easily drying oils such as turpentine or naphtha followed by heating and "burning in" the coating by a suitable hot air blast. App. is described.

Protective coatings on metal pipes, etc. H. KORN A-G. Brit. 338,983, March 11, 1929. The material to be coated is preliminarily heated to the m. p. of a fusible intermediate artificial resin such as one formed from phenol or urea and formaldehyde and such a material is applied (without a solvent) in pulverulent form (either as a dust or in suspension in water or other non-solvent liquid or by rubbing on as a solid block). During hardening the coated surface may be treated with a solvent or solvent vapor and fillers also may be added.

Reducing the viscosity of nitrocellulose solutions. WM. C. WILSON (to John S. Stokes) U. S. 1,703,018, March 10. A phenolic condensation product contg. an active methylene compd. such as $(CH_2)_2N_2$, is added to a nitrocellulose liquor to lower its viscosity.

Cellulose-covered wire gauze. KALL & CO. A-G. (Adolf Dillitz, inventor) Ger. 515,531, June 28, 1928. The gauze is dipped alternately into a cellulose ester solu. and a weak acting swelling agent. The product may be coated with a thin layer of dil. nitrocellulose. In the example, a wire meshed wire gauze is dipped into acetyl cellulose lacquer and, when dry, into an aq. alk. as swelling agent. The operations are repeated until the required thickness is attained. Cells or holes may be used instead of any ale. The finished product may be rendered weather proof by a coating of 2-3% triethyl phosphate.

Linoleum. H. CLAYTON (to Armstrong Cork Co.). Brit. 338,005, Oct. 15, 1928. The oxidizable components of a linoleum cement contg. a siccatve oil are first substantially completely oxidized (suitably as described in Brit. 305,030, C. A. 23, 4817) and a coagulant is added to the cement (about 1% of H_2O_2 , being preferred as coagulant), and the material is mixed with suitable fillers. Numerous details of compn. and procedure are described. Brit. 338,078 specifies the use of a "delayed action" coagulant such as dehydrated boric acid in the form of metaboric or pyroboric acid or boric anhydride, or urea.

Resins from cellulose material. VCHENKO-OLIVNIY ZAVOD MOSKOVSKOY TEKHNIKI, ZINIKOVY PROMISHLENNOSTI "TEKHNIKAVOD" Russ. appl. 50,420, Oct. 15, 1929. Sunflower husks are heated under pressure with water in an aq. soln. of alkali followed by hydrolysis with mineral acids and extr. by the usual methods.

Artificial resin compositions for lacquers, etc. I. G. FARMINSKY A-G. Brit. 339,274, Sept. 3, 1929. Homogeneous products are prepd. by mixing, in soln. or otherwise, a polymerization product of a vinyl ester such as polyvinyl chloride with the condensation product of a phenol derlv. and aldehyde (such as the product from anisole and CH_2O) or the like. An example is given for making a lacquer with CH_2 as solvent.

Molding synthetic resins. A. V. KRITER. Brit. 338,015, Aug. 28, 1929. Decorative articles are made from synthetic resins by placing lace or other fabric (without preliminary impregnation) between a covering sheet and backing of initial condensation product and heating the materials under pressure. Cf. C. A. 25, 610.

Synthetic resins such as urea-formaldehyde resins. I. G. FARMINSKY A-G. Brit. 338,037, July 29, 1929. Condensation of reacting materials such as urea or thiourea and CH_2O for obtaining of non-inflammable products is effected in the presence of org. solvents formed at least mainly of halogen substituted aliphatic aces, such as mono- and dibromobenzyl aces, and chlorobenzyl aces, and preferably in the presence of an acid reacting substance such as oxalic acid or $HClAc$ (the reaction being accelerated by heating). Numerous examples with details of procedure are given, for producing products suitable for making lacquers, molded articles, etc.

Synthetic resins. J. M. WILSON. Brit. 338,545, July 10, 1929. Polymerization of comarone, indene and their homologs is effected by using as the polymerizing agent

a mixt. of an adsorbent earth such as fuller's earth and an acid sulfate of a trivalent metal such as acid ferric sulfate. Various details of procedure are described.

Synthetic resins. COMPAGNIE FRANÇAISE POUR L'EXPLOITATION DES PROCÉDÉS THOMSON HOUSTON. Fr 37,231, July 19, 1929. Addn. to 654,175 (C. A. 23, 3822). The production of artificial resins of the alkyl kind combined with drying oils is described. Thus, a mixt. of glycerol and phthalic anhydride is heated progressively to 200°, a mixt. of phthalic anhydride and one or more fatty acids derived from drying oils (china wood oil, linseed oil or Perilla oil) is added and the heating is continued until the foam disappears, then the desired amt. of drying oil is added, e. g., 10-50% of the wt. of the resin. It is preferable to incorporate the drying oil before complete condensation has taken place. Instead of acids from a drying oil, one or more oxidizable unsatd. acids such as oleic acid, stearic acid or linoleic acid may be used. In another example an unsatd. fatty acid, e. g., an acid derived from linseed oil, is heated with glycerol to 200-250°, linseed oil is then added until a clear soln. is obtained. The soln. is mixed with phthalic anhydride and heated further. The compns. of resin and oil are sol. in acetone, AcOBu, BuOH, triacetin, etc., and may be used as a varnish, etc., or incorporated with fillers for making molded objects.

27—FATS, FATTY OILS, WAXES AND SOAPS

R. SCHERUBEL

Little-known Brazilian oil seeds. P. W. FAIRIE. *Senssieder-Ztg* 58, 131-2 (1931), cf. C. A. 24, 1238.—F. tabulates the consts. of 27 varieties of little-known Brazilian oil seeds, giving the source, popular and botanical names and the time of ripening.

The fatty oil of seeds of *Colchicum autumnale* L. G. GAAL. *Magyar Gyógyszerész-tud. Társaság Értékeit* 6, 149-67 (1930).—The consts. of the oil are dis. 0.9104-5, n_D^{20} 1.4723, acid no. 14.70, sapon. no. 190.0-98, acetyl no. 12.3, Reichert Meissl no. 2.41, Polenske no. 0.32, Hehner no. 91.6-92.1, I.-Br. no. 92.47, satd. fatty acids 12.73%, glycerin 10.8%, mol. wt. 269.5. The fatty acids contained 2.1% raw phytosterin. The presence of arachidic acid could not be demonstrated. Daturic acid was found in the oil.

S. S. DE GUKALY

Experiments with the shipment of some tropical products containing oil in pressed bales. W. W. SPOON. *Ber. Afdel. Handelsmuseum Ver. Koloniaal Inst.* No 56, 23 pp (1930).—With copra the results were satisfactory, there being no loss of oil and only a small increase in free fatty acids. The results with earth nuts and rubber-seed kernels were very unsatisfactory, as both had suffered from the pressing and the percentage of free fatty acids had increased considerably.

J. C. JURKENS

New compounds produced during the hydrogenation of fish oils. II. The formation of the higher alcohols. SEIICHI UENO AND RICHICHO YAMASAKI. *J. Soc. Chem. Ind., Japan* 34, Suppl. binding 35-7 (1931), cf. C. A. 25, 612.—One kg. of unsapon. matter was washed with methanol to sep. hydroxy compds. from hydrocarbons. After evapn. of the methanol the residue was taken up with ether, treated with $\text{Na}_2\text{S}_2\text{O}_8$ and phenyl hydrazine to remove aldehydes and ketones and then washed, dried and evapd. The higher alcs. obtained were fractionated by distn. and the fact was confirmed that the unsapon. volatile matter produced during hydrogenation contains iso-aliphatic C_8 , C_{11} , and C_{14} primary alcs. of the satd. and olefin series together with hydrocarbons of the iso-paraffin and olefin series.

E. SCHERUBEL

Catalytic fat hardening. I. Progressive hardening decreases the speed of hydrogenation. II. Selective hydrogenation of fish oils. L. UBBELOHDE AND H. SCHÖNFELD. *Z. angew. Chem* 44, 184-7 (1931).—The decreasing speed in progressive hydrogenation of oils is not due to the catalyst, since it can be used over several times, each time it shows a high speed at the start and slows up toward the end. Nor is it due to the concn. in the oil of the double C bond, since linseed oil and fish oils are hardened less rapidly than cottonseed oil. It is due to the increasing concn. of satd. acids and to the manner of hydrogenating. By detg. the mol. bromides, formed by highly unsatd. acids, at the different stages during fish-oil hardening, U. and S. confirmed the fact that clupanodonic acid can be reduced completely to less unsatd. acid without forming stearin, by means of a gentle hydrogenation, reduction of the amt. of Ni, temp. and agitation. The fish-oil odor disappeared at I no. 95 when rapidly hardened and the clupanodonic acid at I no. 80-90, slow hardening (45-50 I. nos. per hr.) showed only 0.5% of the highly unsatd. clupanodonic acid remained at I no. 120.

P. ESCHER

Solubility of fats in various solvents. Solubility of Tsubaki oil and rape oil in ethyl alcohol. K. HASIN. *J. Soc. Chem. Ind., Japan* 34, Suppl. binding 64-6, 66-7 (1931).—The soly of refined and bleached Tsubaki oil (*Camellia*, family Theaceae) as well as that of rape oil was detd. in various concns of ethyl alc. by heating known amts. of oil and alc. of different concns in closed tubes and observing the temp. at which both layers turned clear. The turbidity temp. was read during slow cooling. The results are tabulated and expressed in graphs. P. ESCHER

An installation for continuous washing with alcohol in refining oils and obtaining a high percentage of fatty acids and their esters. ORTO KRENS. *Chem. App.* 18, 15-6, 49-51 (1931).—A system is described for washing fatty acids from oils with alc. The alc. removes the acids and is recovered in a rectifying column. M. C. ROGERS

Intramolecular rearrangements when hydrogenating esters of fatty acids with one double bond. II. A. STEGER AND J. W. SCHEFFERS. *Chem. Umschau Fette, Öle, Wachs u. Harze* 38, 61-6 (1931), cf. *C. A.* 25, 2315.—A Δ^8 -elaidic, a Δ^8 -oleic and a Δ^8 -oleic Et ester were partially hydrogenated with Ni at 180° to an I no. of 55, at which point a max. of solid unsatd. acids is formed. The products were analyzed for their solid and liquid acids by Twitchell's alc. method, sep'd into their individual acids by fractional crystn. in different solvents, and the constitution of each acid was detd. by ozonization and analysis of the ozonides for m. p., proximate compn. and sapon no. The double bonds in Δ^8 -elaidic and Δ^8 -oleic acids are shifted during hydrogenation from Δ^8 toward both sides to Δ^{10} and Δ^6 , a greater amt. being formed of the former than of the latter, the elaidic acid forms less of the Δ^{10} and Δ^6 -elaidic acids than oleic acid, and thus it may be assumed that the principal shifting of the double bond occurs in the liquid oleic acid and that the first-formed Δ^6 and Δ^{10} -oleic acids are later converted into Δ^8 and Δ^{10} -solid elaidic acids. The Δ^8 -oleic Et ester shifts its double Δ^8 bond also to its neighboring C atoms, forming Δ^6 and Δ^{10} -oleic acids and possibly also the corresponding Δ^6 and Δ^{10} -elaidic acids. No conclusions could be drawn with respect to the manner in which the bonds of solid unsatd. acids change during hydrogenation, since hardening was stopped at the point of their max. formation. P. ESCHER

The spontaneous ignition of commercial oleic acid. PAUL ERASMUS. *Allgem. Öl-Fettsig.* 27, 809-12, 345-8, 367-8, 408-9 (1930).—Oleic acid manuf'd. from bone fat contained no highly unsatd. acids. It was very susceptible to spontaneous ignition in the Mackey app. When the anhydrides present were removed it was no longer susceptible to spontaneous ignition. The removal of the anhydrides also prevented the darkening of the oleic acid on atm. oxidation. Addn. of the isolated anhydrides to an oleic acid which was not spontaneously ignitable caused it to take on this property. The substance causing the spontaneous ignition was identified as γ -stearyl lactone. This substance on oxidation yields ketonic acid, which in turn gives up O_2 easily to form the lactone again. The addn. of 0.05% of γ -stearyl lactone to oleic acid causes it to become spontaneously ignitable. The presence of Fe was found necessary to bring about this result. Removal of Fe from a spontaneously ignitable oleic acid caused it to lose this property. Addns. of small quantities of Fe oleate to an oleic acid not spontaneously ignitable did not show any effect. When larger amts. (0.5%) were added it became spontaneously ignitable. Conclusion: The lactones are O_2 carriers, requiring the presence of Fe as catalyst, and are a cause of the spontaneous ignition of oleic acid. The addn. of 1% linolic and linolenic acids (lactone free) to an oleic acid not spontaneously ignitable caused it to become so. W. F. BOLLERS

Oleiferous *Allanblackia*. Chemical composition of *Allanblackia floribunda* Oliv. J. PIERAERTS AND SERGE VLASSOV. *Mat. grasses* 22, 8975-7 (1930).—Addnl. analysis has been made supporting previous conclusions that this plant is the cheapest and most satisfactory raw material for the manuf. of pure or technical stearic acid. P. T.

Thermodynamics in the soap industry. K. LÖPFL. *Seifensieder-Ztg.* 58, 127-8, 152-3 (1931).—L. discusses the calcn. of the total calories required for boiling and drying soaps. Spray drying shows simplicity of app. and least heat consumption. P. E.

Bleaching earths (ECKERT) 18. Germicidal efficiency of soaps and of mixtures of soaps with NaOH or with phenols (SCHAEFFER, TILLEY) 11C. The influence of soaps on the germicidal properties of certain mercurial compounds (HAMPTL) 11C. Soaps, creams, etc. (Fr. pat. 695,218) 17. Destructive hydrogenation [of fatty oils] (Brit. pat. 339,048) 21. Complex organic peroxides [for bleaching soaps, oils, fats, waxes, etc.] (Brit. pat. 339,336) 12. Mixtures of hydroxy fatty acids (Ger. pat. 518,390) 10. Substitutes for sugars in [filling transparent soaps] (Ger. pat. 518,196) 18.

HUSSON, M.: Les huiles d'olives. Algiers: Impr. algérienne 14 pp.

MARTIN, GEOFFREY. *The Modern Soap and Detergent Industry*. Vol. I. *Theory and Practice of Soap Making*. 2nd ed., revised and enlarged. London: Crosby Lockwood 36s., net

Preparation for splitting fats. G S PETROV. Russ appl 62,843, Jan. 23, 1930. Addn. to pat 18,855. Fat-splitting agents are prepd by aullonation of highly hydrogenated fats or latty acids, excluding hydrogenated castor oil.

Oils and fats. PETER FRIESENIHAIN. Ger 515,769, May 28, 1925. Addn. to 365,160. The hydrated phenols of 365,160 for prepg emulsions or soins of oils and fats are supplemented or replaced by aics formed by the hydrogenation of unsatd hydrocarbons obtained from the cracking of acid resins or from the residual pitch of aliphatic raw material distns.

The concentration or drying of masses containing fatty oils. L. AUER. Hung 101,164, May 28, 1925. Masses are coned or dried in an atm. contg less O than does air.

Preventing oxidation of oils, etc. COMPAGNIE FRANÇAISE POUR L'EXPLOITATION DES PROCÉDÉS THOMSON-HOUSTON. Fr 695,119, May 6, 1930. The oxidation of oils, fats, waxes, etc., is prevented by adding 0.5-1% of phenyl α naphthylamine.

Extracting oil from blubber, meat, bones, tongues and other parts of whales, etc. C O JOHNSON. Brit 339,305, Aug 31, 1929. In a wet rendering process, the materials are submerged in water which is kept boiling under pressure during the digestion by introduction of steam. App and various details of procedure are described.

Apparatus for extracting oil from olives, etc. SOC. JEANJEAN ET CASAS. Fr 37,352, Aug 9, 1929. Addn. to 654,217 (C. A. 23, 3824).

Reaction products of cashew nut shell oil. THE HARVEL CORP. Ger. 518,424, June 4, 1927. See Brit. 272,510 (C. A. 22, 1863) and 275,574 (C. A. 22, 2285).

Soap. JEAN GNAEDINGER. Fr 37,464, Sept 27, 1929. Addn. to 665,425 (C. A. 24, 984). The α -abietic acid is converted to the β form which is more stable before combining it, by the aid of catalysts, with unsatd. oils of the fatty series to form a soap.

Soaps. I. G. FARBENING A-G (Otto Ambros and Anne Nies Harteneck, inventors). Ger 518,019 Mar 13, 1928. Products yielding stable foams and emulsions are prepd by adding soft or hard fish roe or prepgs. thereof to soaps or saponaceous materials. A suitable addnl substance may be prepd by removing fat from soft fish roe by means of an org. solvent and then drying, or the active material may be extd from the roe with dil. alkali.

Dry-cleaning soap. LEO R. HUBBARD (to Camille Dreyfus). Can 309,361, Mar 10, 1931. A detergent adapted for the dry cleaning of fabrics comprises a mixt of oleic acid 1, cyclohexanol 1, CCl_4 1, ammonia of (30° B \acute{e}) $\frac{1}{2}$ and water $\frac{1}{2}$ part.

28—SUGAR, STARCH AND GUMS

J. E. DALE

Quality of the mineral fuel received by sugar factories for the campaign 1929-30. I. B. MINTZ. *Nauk Zapiski Tekhnolos Prom* 10, 225-35 (1930).—A report.

V. E. BAIKOV

Movement of nitrogenous substances in sugar manufacture during the campaign of 1929-30 at the Usinskii sugar factory. B. A. LYASKO. *Nauk Zapiski Tekhnolos Prom* 10, 23-38 (1930).—From a no. of analyses made during the campaign I. shows the amt. of noxious N represents about $\frac{1}{2}$ of the total N of the beet. At the diffusion battery about $\frac{1}{2}$ of the total N, $\frac{1}{2}$ of the albuminum N and $\frac{1}{2}$ of the noxious N are eliminated, while ammonia and amide N go almost entirely into the diffusion juice. 96.5% of N in sirup is of the noxious type. Ninety % of the noxious N passes into the final molasses with 10% eliminated in the diffusion battery. Ten tables and 2 diagrams are shown.

V. E. BAIKOV

Experiments in the harvesting of burned cane. III. P. O. J. 213 cane. H. H. DODDS AND P. FOWLER. *Proc S African Sugar Tech Assoc* 1930, 122-32 (1931), cf C. A. 24, 3125.—It is shown in a no. of tables and graphs that the variety P. O. J. 213, like the Uba cane, deteriorates more slowly after being burned than when it is not fired. The burned cane keeps better if left standing than after it is cut. The differences between burned and unburned canes are even greater than in the case of Uba, and unburned cane must be taken to the mill without delay, while burned cane may be kept for about a week. But in fields which have been burned over, the supply of org. matter must be maintained by fertilizing.

F. W. ZERAN

Determination of small proportions of invert sugar in raw sugars. LEWIS GYNON AND J HENRY LAMP *J Soc Chem Ind* 50, 85-61(1931), cf *C. A.* 17, 1400, 2091; t8, 763—The original tables were calcd only for products contg not less than 0.3% invert sugar. A new table is presented for a range of 0-0.461% invert sugar. Twenty-five g of the sugar plus 0.1 g invert sugar in the form of the standard soln is dissolved to 100 ml. This soln is titrated against 10 ml Fehling soln by the standard method previously described. It is unnecessary to clarify with $\text{Pb}(\text{OAc})_2$ or to remove lime salts with $\text{K}_2\text{C}_2\text{O}_4$, because the same results are obtained without their use.

F. W. ZERBAN

Measurement of color in solutions of white sugars. J. A. AMBLER AND S. BYALL *Ind Eng Chem, Anal Ed* 3, 135-6(1931)—The measurement of the colors of white sugars was placed on a reproducible, accurate basis by matching the colors of water solns of the sugar against a standard made by dissolving 10 g of Balch's No. 5 maple sirup color standard soln (*C. A.* 24, 2630) with water to 100 cc. For the detn of color of hard candies made from the sugars the candies were dissolved in water and the colors of the solns matched against a No. 9 maple sirup standard soln, dild in the same way as the No. 5. The color of the barley candy is not related to that of the original white sugar.

F. W. ZERBAN

Impurities in white sugars. I. Determination of phosphorus. S. BYALL AND J. A. AMBLER *Ind Eng Chem, Anal Ed* 3, 136-7(1931)—The importance of traces of impurities in sugar necessitates analytical methods of greater sensitivity. Total P_2O_5 is detd in the HNO_3 soln of the residue obtained by ashing in the presence of Na_2CO_3 , inorg P_2O_5 is detd directly in a water soln of the sugar, and org P_2O_5 is found by difference. The Briggs modification (*C. A.* 16, 3197) of the Bell and Dorsey ceruleo-molybdate method (*C. A.* 14, 3685) detects as little as 0.3 p. p. m. of P_2O_5 in a 10 g sample. Analytical results on 26 samples of white sugars show up to 1.4 p. p. m. of inorg, and up to 45.6 p. p. m. of org, P_2O_5 .

F. W. ZERBAN

Recovering of ammonia in beet-sugar manufacture. A. I. VOSTOKOV AND M. P. KOTLYARENKO *Nauk Zapiski Tsvukrovoi Prom* 10, 211-9(1930)—The methods for recovering NH_3 from evapd water from beet juice are described. The product is a high-grade fertilizer.

V. E. BATKOV

The German report on the Oxford sugar method. OSKARI ROUTALA *Acta Chemica Fennica* 3, 94-5(1930)—If sugar beets were not treated immediately after picking the yield of sugar was low and there was considerable inversion, especially if the beets were stored in a warm place. The yield was also low if the beets were frozen. The Oxford sugar method consists in the rapid picking, cutting and drying of the sugar beets to preserve the sugar content for later use. The beets were cut into very small pieces and dried in a 3 compartment oven at 60-70° for 1 hr. The dry beets had 60-70% sugar, compared to 7-8% in the undried beets. Very little inversion was found. The drying affected the protein and protoplasmic content of the beets in such a way that little of these substances was found in the sugar solns. The sugar from the dried beets was extd very readily by water. The soln turned black when 0.2-0.3% of Ca as oxide, pectate, oxalate or phosphate was added. Addn of activated charcoal and centrifuging gave a colorless sugar soln. The molasses obtained was sweeter than ordinarily, because of a higher invert-sugar content. The German investigation comm. reported that the procedure is as yet not industrially profitable since the change to the new equipment is costly.

S. A. KARJALA

Experiments on raw sugar with Russian activated carbon. M. P. KOTLYARENKO AND N. K. KONIBOLOZKII *Nauk Zapiski Tsvukrovoi Prom* 10, 102-11(1930)—The decolorizing power of the Russian activated C is higher than that of Norit-Standard, Amsterdam, 1928. The adsorption of colloids and ash by Norit is smaller than that by Russian C. The colloids and ash are not retained by the C and are almost entirely washed back into the thick juice. The wash (sweet) water must be given a preliminary treatment to eliminate the ash and colloids before it is mixed with a sirup. The max. regenerating effect of activated C was obtained by thermal and HCl treatment.

V. E. BATKOV

The quality of the sweet (wash) water from the bone-black filters. M. I. NAKHMANOVICH AND I. P. ZELIKMAN *Nauk Zapiski Tsvukrovoi Prom* 10, 51-74(1930).—From a no. of expts the following conclusions can be drawn: With decrease of the density, the proportion of non sugars increases. The sweet water washes out from the chars the coloring matters adsorbed in the early stages of the filtration. Washing should be stopped when the sweet water decreases to 2.5 Brix (1.5-1.6% of sugar).

V. E. BATKOV

North installation at the Khutor-Mikhailovskii sugar refinery. G. I. ZHURIK. *Nauk Zapiski Trukrovot Prom* 10, 86-102(1930) — A description V. E. BAIKOV

Regeneration of the bone chars. N. S. VOLKOV. *Nauk Zapiski Trukrovot Prom* 10, 294-307(1930) — Regeneration of bone chars can be divided into 2 parts: (1) boiling for a certain period of time in the same vol of water, and (2) washing with clean water for a limited time. With poor quality of water, the time of boiling and amt of water should be increased. While the water dissolves mineral salts and organic matters, the bone black adsorbs coloring matters from the water. If the temp of the wash water is decreased from 90-95° to 55-60°, the amt of water must be increased 30-40%. Washing with condensed water has the advantage of decreasing the amt of steam. The chem control of the regeneration consists in checking and comparing the mineral residue of the wash waters before and after washings.

V. E. BAIKOV

Prospects and advantages of rotary filters in sulfitation factories. W. H. FOSTER. *Proc S African Sugar Tech Assoc* 1930, 71-6(1931) — The sucrose (2%) lost in filter cake in Natal is higher than in other countries. Higher diam of settlings is not economical. Washing in the presses gives indifferent results, and furnishes a sloppy cake. The use of filter aids in sufficient quantity is too expensive. Double filtration through the presses reduces the sucrose loss in filter cake, but also the purity of the filtrate and the boiling house efficiency. Double settling is being practiced, but it requires too much steam for evapn. The Maus continuous rotary vacuum filter has been tried in various ways. The 2nd operation in double filtration can be carried out more economically with the rotary filter, the machine being particularly suitable to this type of work. The low purity of the 2nd filtrate is an undesirable feature of the process. The solid content of the settlings is rather low for the successful direct operation of the rotary filter. Should a thickened sludge be possible at low cost, the filter offers considerable advantages over the filter press as regards economy in labor and cloth, increased sugar recovery and greater cleanliness. The floor space required is much greater than in the case of filter presses.

T. W. ZERBAN

Metafiltration in the sugar industry. ANON. *Intern Sugar J*, 33, 18-20(1931). — It is a well-established fact that the filtration employed in the carbonatation process is responsible for the increased yield, and the better quality of the sugar as compared with the defecation or the sulfitation process. Filtration app., whether of cloth, sand, wire gauze or rings as in the case of the Metafilter, act simply and solely as a support to the solid matter which it retains. The filtering elements of the Metafilter are made up of thin brass rings $\frac{1}{8}$ " outside and $\frac{1}{16}$ " internal diam. The rings are piled one upon another to form a tube so that a 1 in. of filtering elements is composed of some 30 rings. Alternate rings carry a light embossing on each face exactly 0.002 in. high, which provides a gap of this thickness between each pair of rings. The rings are strung on a slotted rod which acts as a support and provides drainage channels, the rings being pressed together by screw pressure from the end. An ingenious aspect is an arrangement to rotate the whole filter body *en bloc*. In operating the Metafilter on filtering juices in a defecation or sulfitation factory it is necessary first to precoat the column of rings with a thin film of kieselguhr. It was found in Java that an addn of 0.5% on the Brix of the heavy liquor being filtered was all that was necessary to obtain a fairly rapid flow. In 60 min., in one run 1000 l of 55-60 Brix juice at a temp of 70° was filtered with a filtering area of 2.2 sq. m., corresponding to a capacity of 7.5 l per sq. m., per hr., a very high rate of filtration.

W. L. OWEN

Loss in polarization of Natal raw sugars on storage. L. BLACKLOCK. *Proc S African Sugar Tech Assoc* 1930, 105-7(1931) — Several cases of excessive deterioration of Natal sugars call for remedial measures. The safety factor should be kept low, by drying if necessary. The sugar should be of uniform grain and not acid. The mill must be kept in sanitary condition. Mixing of high grade with low grade sugars should be avoided. Storage and transport conditions should be as perfect as possible, low humidity and low temp being essential.

F. W. ZERBAN

Palmyra palm and its uses in gur manufacture. M. R. GOKARN. *Poona Agr. Coll Mag* 21, 241-52(1930) — The manu. of gur (crude sugar) from the juice obtained by tapping the immature nuts of the palmyra palm (*Borassus flabellifer*) is described. The best yields of juice are obtained from the female spathes, the av. yield per tree ranging from 60 to 100 gallons per season. Saccharimeter readings on several samples of palmyra juice ranged from 11° to 13° Brix as compared with 13.5° to 18° Brix for coconut juice. There was no marked difference in the saccharimeter readings on the juice from the male and female palmyra spathes. Two samples of the crude sugar contained 2.10 and 2.61% glucose and 86.68 and 83.51% sucrose, resp.

K. D. JACOB

New sources of Inulin. H. VANOVERKY and R. M. KINSMARKY, *J. Agr. Chem. Soc.*, 53, 1507, 601 (1931).—There are two sources of inulin (1) are discussed: Camassia roots (*Camassia esculenta* (Nutt.) Coville, (*Camassia esculenta* (Nutt.) Coville), of which 200 g. gives out 1 g. of I. *Chikusio nuttallii* (Walt.) Wats., a variety of wild onion, 200 g. of which gives 25 g. I. and *Polygonum eriantherum* (Nutt.) Nutt., belonging to the family Compositae, 100 g. of which gives 52 g. I. The ap. rotation of I from the two sources ranged from -2° to -31° , all products were sol. in H₂O and dil. HCl. In the light of this work and that of Chavastion (*J. Pharm. Chem.*, 2, 241 (1931), it appears that there is no correlation between the ap. rotation and the body of I. No claim is made that any of the products described was a chemically pure inulinol.

Measurement of consistency of starch solutions, J. C. RUSSELL, *Ind. Eng. Chem., Anal. Ed.* 3, 125-126 (1931). The swell of starches is measured by pointing the starch suspension prepared by a standardized procedure, into water in a cylinder, allowing it to stand and measuring the vol. occupied by the starch granules on the next day. It was found that the product of concn. and swell plotted against the viscosity, furnishes a curve which may be used to evaluate starches without a viscosity determination. The method gives best results with starches showing large swell like potato, tapioca and arrowroot starch.

Potash fertilizer for sugar beets (WATSON) 15. The change in x-ray spectrum of inulin on swelling in water (KATO, FUKUNAGA) 2.

Sugar Abstracts (New Journal). Published by the Palmer Publishing Corp. of New York, 161 Waverly Place, New York City. Appears as a part of *Food Abstracts* and also as a separate journal. Vol. 1, No. 1 appeared in Jan., 1931. Price, \$5 per year.

Sugar, INTERNATIONAL PATENT DEVELOPMENT CO. Bull. 240, 1931, Sept. 10, 1931. Sub. 11, 1931, 1931 (C. & S. 24, 415).

Calandria apparatus for evaporating sugar solutions, etc., JAMES HAMILL, JOHN V. TANNIKYAN and GEORGE W. CONNOR. U. S. 1,705,094, March 10. Structural features.

29 LEATHER AND GLUE

ALLEN PROVER

Leather, sanitation and colloid chemistry, JOHN A. WILSON, *Ind. Eng. Chem.*, 23, 417-421 (1931).—An abstract.

New Zealand made leather, P. WHITE, *New Zealand J. Sci. Ind.* 1930, 251-7; *J. Intern. Soc. Leather Trades Chem.* 14, 561.—Results of chemical, water absorption and wear tests are given.

Gas poisoning in the leather industry, A. GUNNINGHAM, *Vestnik Khimicheskoy Prom. i Torgov. (Vestnik of Leather Ind. and Trade)* 1930, No. 7/8, 411-4; *J. Intern. Soc. Leather Trades Chem.* 14, 561-90. A fatal poisoning occurring when workmen entered an empty vat was traced to HCN, produced from amygdalin of pine bark by enzymic action. Under certain manipulative conditions this gas may form a layer in the bottom of the vat when the liquid is drawn away. This may be dispelled by a water jet. Safety rules are given.

Stretching tests for leather, C. A. HEAVIS, *Bull. U.S. Dept. Agr., Ind. and Coll. Res. Council* 8, 1-6 (1930); *J. Intern. Soc. Leather Trades Chem.* 14, 565.—Samples cut from different positions in the same skin show such variations that it is dangerous to draw conclusions even if from a large no. of skins. A stretching test is now that required for rupture causes permanent elongation. The relation between elongation and contraction in width and thickness is not the same as that obtained with hydrostatic stress. Simple rupture and tearing rupture in a direction not at right angles to the strain are distinguished.

Estimation of fat in leather, P. KOTVINICH, *Vestnik Khimicheskoy Prom. i Torgov. (Vestnik of Leather Ind. and Trade)* 1930, No. 4, 147; *J. Intern. Soc. Leather Trades Chem.* 14, 574.—Skin with poly ether for tanning (displacement), without previous drying of the leather, is advised.

Determination of water in vegetable leather, R. PARANAY ENNEAANNI, J. M. COMPTON, *J. Intern. Soc. Leather Trades Chem.* 15, 125-130 (1931). Measurements in the oven drying method are studied. Factors considered are those of heating, pressure and time (various at 95° F. at 110° F.), and grease content of the leather. Many of the findings

are contradictory and it is concluded that the accuracy of the detn. of water in vegetable leather by the oven method cannot be guaranteed no matter what the exptl. conditions, unless the grease content of the leather is extremely low. Regain in wt. of oven-dried samples stored over corrod. H_2SO_4 is ascribed to adsorption of air. H. B. MERRILL.

Irgiten and its use in dyeing chrome leather. J. G. GRIGY. *Boll. ufficiale staz. sper. ind. pell. mat. concen.* 8, 154 (1930), *J. Intern. Soc. Leather Trades Chem.* 14, 435.—Methods of application are described. H. B. MERRILL.

Waterproof finishes [lacquer finishes for leather]. D. WOODROFFE. *Leather Trades Review* 1930, *J. Intern. Soc. Leather Trades Chem.* 14, 318.—A discussion. H. B. MERRILL.

Hydration of animal skin by the volume change method. IV. Effect of various factors on the hydration of calf skin. EDWIN R. THREIS AND P. T. BENTON. *Ind. Eng. Chem.* 23, 267-70 (1931); *C. A.* 23, 2319.—The dilatometer method has been improved so as to measure hydrations quantitatively in ml. Postmortem changes reduce hydration and hasten the onset of degenerative dehydration. The onset of degenerative dehydration is prevented by salts and antiseptics, which indicates that bacterial action is responsible. Hydration, though in reduced amt., takes place in satd. solns. of salts, indicating that the wt. loss accompanying curing is due to removal of free rather than bound water. Hydration in satd. salt and alkali solns. follows the order $LiCl$ ($NaCl + Na_2CO_3$) $< MgSO_4 < NaCl < 0.05 N NaOH < Ca(OH)_2$, + excess lime $< H_2O$, $Ca(OH)_2$. With dry-cured skin, $NaCl$ in low concn. (2-4%) hastens the attainment of max. hydration but also hastens the onset of degenerative dehydration. In moderate concns. (8-10%) it delays the attainment of max. hydration without diminishing its extent, and inhibits dehydration. In high concns. (15-20%) it decreases the extent of hydration. Hydration increases with the ratio of water to the skin in the total exptl. system, and decreases with increasing temp. (again contrary to the effect on swelling). Increase in sp. gr. by the displacement method can be used to measure hydration. H. B. MERRILL.

Physicochemical differences of hide powder by sex. T. TADOKORO AND K. YOSHIMURA. *J. Soc. Chem. Ind., Japan* 34, Suppl. binding 37-8 (1931).—Hide powder prepared from bull hide showed the following differences from that prep. from cow hide: greater "acid producing power" in neutral salt soln., greater tannin absorptive power, less I_2 absorptive power, less swelling capacity in salt solns. (no difference if hide powders are purified). H. B. MERRILL.

Microscopic study of the effects of cold temperatures upon skins and hides. FRED O. FLAHERTY AND WM. T. RODDY. *J. Am. Leather Chem. Assoc.* 26, 172-80 (1931).—No damage was observed that could be attributed directly or indirectly to low or sub-freezing temp. Keeping skins or hides at low temp. has a tendency to dehydrate them, which also aids in their preservation. Frozen skins should never be flexed and should be allowed to thaw before soaking. H. B. MERRILL.

Columbian hides. V. A. LATHOM SHARP. *Leather Trades Review* 1930, ES9-92, *J. Intern. Soc. Leather Trades Chem.* 14, 587.—Methods of cure and preservation are described. H. B. MERRILL.

Investigations on animal skin. GEORG GRASSER AND HIROSHI OHOKI. *J. Faculty Agr. Hokkaido Imp. Univ.* 27, Pt. 2, 263-74 (1930).—The time of hydrolysis of skins of different species in solns. of acids, bases and salts is tabulated. H. B. MERRILL.

Investigations on gelatin and skin. GEORG GRASSER AND HIROSHI OHOKI. *J. Faculty Agr. Hokkaido Imp. Univ.* 27, Pt. 2, 275-84 (1930).—The peptizing effects of numerous salts are listed. H. B. MERRILL.

Cultivation and tannin content of *Caesalpinia coriaria* (divi-divi) in [former] German E. Africa. T. MARX. *Tropen* 33, 100-5 (1930), *J. Intern. Soc. Leather Trades Chem.* 14, 587-8.—The tannin content is fair after 9 years' growth and excellent after 15 yrs. The percentage of tannin compares favorably with that of Indian divi-divi. H. B. M.

Problem of tannery waste. D. JOHANN LEYD. *Leather World* 22, 1102-5 (1930).—Products are enumerated and current practice for their disposal is discussed. H. B. MERRILL.

Tanning materials and their characterization. II. Organic tanning materials. GEORG GRASSER AND HIROSHI OHOKI. *J. Faculty Agr. Hokkaido Imp. Univ.* 27, Pt. 2, 242-61 (1930); *C. A.* 24, 4951.—Contraction temp. and time required for complete hydrolysis on boiling are tabulated for hide and gelatin tanned with a large no. of tanning materials, including natural tannins, syntans, aldehydes, phenols and their derivs. The effect of added salts and of solvent (in case of aldehydes) is studied. No general relationships are found and no conclusions drawn. H. B. MERRILL.

Analytical control of sulfonated oils used in the tannery. M. GERONAZZO. *Boll. ufficiale staz. sper. ind. pell. mat. concen.* 8, 142-50 (1930), *J. Intern. Soc. Leather*

Trades Chem 14, 373; cf. *C. A.* 22, 2485 —The procedure devised for sulfonated castor oil has been applied to sulfonated animal and fish oils. The results are sufficiently precise for technical purposes. H. B. MERRILL

Combined working of two tanning materials on gelatin and animal's skin. Chemical investigation of combination tannage. GEORG GRASSER AND MASATAKE ICHISE *J. Faculty Agr. Hokkaido Imp. Univ.* 27, Pt. 2, 319-48 (1930) —The contraction temp. of the skin and m. p. of gelatin tanned successively with 2 tanning materials are tabulated. H. B. MERRILL

Mangrove as a tanning material. T. A. BUCKLEY *Malayan Forest Records* 1929, No. 7, *J. Intern. Soc. Leather Trades Chem.* 14, 277-8 —Analyses show the variations in tannin content in bark from different species, different trees and different parts of the same tree. Great variations exist; the average compares unfavorably with Madagascar and E. Africa mangrove. *Rhizophora mucronata* has very high tannin content (28% of moist bark); the more plentiful *R. contigata* contains only about 16%. *Carapa obata* is the species richest in tannin but its extracts froth badly. Bark should be extracted as soon as stripped; there is a gain in tannin on drying but this is offset by objectionable darkening of color. The tannin content tends to increase from the bottom to the top of the tree, and is highest in older, bigger trees. H. B. MERRILL

The Pavlovitch alkaline tanning method for the preparation of sole leather. L. KLEIBLOUKIS *Lederfach Rundschau* 1930, 81, *J. Intern. Soc. Leather Trades Chem.* 14, 574 —The leather is tanned first with quebracho at $pH = 8 - 7.5$ for 6 days, the concn. and temp. being gradually increased, then with oakwood extract at $pH = 3.8 - 4.0$. Analyses and wear tests show the leather so made to be normal. H. B. MERRILL

Tanning small fur skins. O. DUJARDINE *Leather Manufacturer* 41, 260-1 (1930) —A description. H. B. MERRILL

Tanning reptile skins. D. WOODROFFER *Leather Trades Review* 1930, 810, *J. Intern. Soc. Leather Trades Chem.* 14, 573. H. B. MERRILL

The theory of chrome tanning in the light of detanning processes. C. STIEPEL *Kunstdünger u. Leim* 28, 15-21 (1931) —Chrome tanning proceeds in 2 stages: beginning in the alk. bath, but the max. destruction of swelling properties of hide does not result until the bath has become acid from pptn. of a basic Cr compd. upon the fibers of the hide. In detanning chrome scrap for glue, 2 products are obtained, one H_2O sol and of normal behavior, the other H_2O insol, so-called "chrome glue" obtained by the magnetite process (cf. Stiepel, *C. A.* 25, 2018), which represents hide substance tanned under acid conditions, more difficult to detan. C. J. SCHOLLENBERGER

Tanno-chemical investigation of chrome salts. GEORG GRASSER AND MASATAKE ICHISE *J. Faculty Agr. Hokkaido Imp. Univ.* 27, Pt. 2, 295-318 (1930) —The contraction temp. of the skin, and m. p. of gelatin tanned with different chrome salts, prepd. under different conditions, are tabulated. H. B. MERRILL

The gel strength of glue from hide, from chrome leather and from bones and its relation to the viscosity and chromium content. O. GERNIGROSS AND H. MENDEL *Kunstdünger u. Leim* 28, 109-16 (1931) —Ten hide glues, 5 bone glues and 13 glues from chrome leather were investigated. Procedures are described in detail. Gel strength of a 12.5% soln. at various temps. was detd. with the Greiner app. Viscosity of a 17.75% soln. at 40° was measured by the Vogel-Ossag app. The pH of a 1% soln., water and ash contents of samples are also given, but methods are not described. Cr_2O_3 in chrome glues varied 0.03-4.20%. Among com. Cr glues, there was no connection between Cr_2O_3 and phys. properties, but 3 samples from leather dechromed in the lab. showed progressive improvement in gel strength and viscosity with Cr removal. Chrome glues (with one exception) and bone glues showed higher gel strength than hide glues of the same viscosity. The relation of gel strength to viscosity was fairly regular among all samples and most regular among hide glues. The pH varied 5.0-7.6, and within this range had no important effect upon viscosity. C. J. SCHOLLENBERGER

Influencing of gelatin-metal salt precipitation by addition of neutral salts. GEORG GRASSER AND HIROSHI OHOKI *J. Faculty Agr. Hokkaido Imp. Univ.* 27, Pt. 2, 285-8 (1930) —The effects of a large no. of Na salts on pptn. of gelatin by salts of weak bases and heavy metals are listed. H. B. MERRILL

Ashing glue and gelatin in the "Effix" muffle furnace. ERNST GOEBEL *Kunstdünger u. Leim* 28, 117-9 (1931) —Tests of a new muffle furnace showed that glue and gelatin samples could be ashed to const. weight at a saving of 50% time and 80% gas as compared with ignition over a Bunsen burner. Comparisons of crucibles of different materials showed that least time and gas were required with Pt, slightly more with "Weta" ware and most with silica and porcelain. With the latter, time and gas consumption were approx. 100% greater than with Pt. C. J. SCHOLLENBERGER

Animal materials for artificial silk, a chemico-technical problem for the glue and gelatin industry. H MENDEL. *Kunstidunger u Leim* 28, 130-1(1931).—If filaments of gelatin could be tanned or fixed while under tension, the tendency of the moils to orient themselves (cf Katz and Gerngross, *C A* 20, 628) should result in a fiber with superior strength, which might compete with artificial silk. Investigation of this or similar problems might well be undertaken by the industry. C J SCHOLLENBERGER

The collagen fiber. I HALLA and R TANDOLF. *Z physik Chem*, Abt. B, 12, 89-92(1931).—X ray spectrograms of collagen fibers from fish skin were identical with those of gelatin under tension. IRANK URBAN

Changes in the use of adhesives. HERMANN STADLINGER. *Kunstidunger u Leim* 28, 8-11(1931).—A general outline is given of the uses of natural adhesives in the past, recent substitution of synthetics and possible new applications for old materials. C J SCHOLLENBERGER

Glue and gelatin in the textile industry (STADLINGER) 25. Refractometric study of chromic acid reduction (GRASSE, OHOKI) 6. Study of the warble and its eradication (GANSSE) 15. Uniting rubber and leather (Brit pat 338,789) 30. Substitutes for sugars in [chrome tanning] (Ger pat. 518,196) 18.

STIASNY, ED. Gerbereichemie. Dresden T. Stemkopf About 600 pp. About M 35, bound, about M 37.

WIENNA, FERDINAND, and WENNA, H. Die Lederfärberei und die Fabrikation des Lackleders. Ein Handbuch für Lederfärber und Lactierer 4th ed, revised and enlarged Leipzig A Hartleben's Verlag 220 pp M 5, bound, M 6

Rotary-drum tanning apparatus. E G WILSON Brit. 339,252, Aug 23, 1929. Structural features

Removing fat from sheepskins. P I PAVLOVICH, N G SHCHEKOLDIN, V I DAVLING, K I ZYABLOV and V F DROZDOV Russ appl 68,138, Apr 11, 1930 Skived sheepskins are treated with caustic solns of "Kontakt" (petroleum sulfonic acids) or its soap

Chrome velvet leather. V S AVTONOMOV Russ. appl. 59,814, Dec. 7, 1929 Chrome leather is treated with warm water followed by an alternating treatment with a soda soln and Cr alum, and then by the usual treating methods.

Leather for driving belts, etc. A MCLELLAN Brit. 338,536, Aug 10, 1929 Leather for belt conveyors or driving belts is rendered resistant to moisture and wear by buffing and treating with a soln formed of gutta serena, Sb sulfide or other cold vulcanizing agent, and CS₂ or CCl₄ or both (which may be applied in successive coats, each of which is allowed to dry before applying the next coat) The material may be subsequently subjected to pressure, and may be preliminarily impregnated with rubber Various details of solns and procedure are described

Dyeing leather. N P SHAYKOV Russ appl 39,491, Jan 26, 1929 An ext is obtained by boiling the bark of the larch tree or pine tree with caustic remaining after a complete extn of tanning material The leather is dyed with this ext with the usual metal mordants by the customary dyeing methods

Dyes for chrome leather. I G FARBERIND A-G Brit 338,930, July 22, 1929. Aromatic N nitroso compds are warmed with aromatic amino-, hydroxy-, or amino-hydroxy-compds (suitably in the dye bath at the time of dyeing chrome leather) Numerous examples are given for producing dyeings of various colors

Monoazo dyes for leather. I G FARBERIND A-G Brit 339,029, Sept. 24, 1929 Dyes giving brown shades on leather are obtained by coupling a diazotized p-nitroaniline negatively substituted in a position to the NH₂ group, or a homolog, with a compd of the diphenylamine series such as diphenylamine or a sulfo or carboxy deriv Examples are given

Artificial horn. I G FARBERIND A-G (O Schmidt, K Seydel and E Meyer, inventors) Ger 482,930, Aug 28, 1926 Casein is worked up by the customary processes after addn of a small quantity of a liquid or solid hydroaromatic or aliphatic-hydroaromatic base, or an aliphatic base contg an OH group, or a salt or deriv of such a base The use of cyclohexylethanolamine is described in an example

Glue. PORA-WERK PAUL SCHROT. Ger 515,868, April 27, 1926 A filling material for animal glue consists of wood dust and such mineral as gypsum or chalk. The filling material is intimately mixed with the glue under pressure

Glue and gelatin from chrome leather cuttings. BRITISH GLUES & CHEMICALS,

LTD. and R. B. DREW *Brit.* 333,584, Aug 17, 1929 Chrome leather cuttings are subjected to percolation with a soln. of an alk. earth such as magnesia, repeatedly and cyclically in one direction at such a rate that the acid extd. from the leather neutralizes sufficient alkali to maintain the desired alkalinity (a pH of 8-8.5 being suitable). A temp. of 105° is suitable for thick cuttings and undesirable sol. matter such as salt may be removed by a preliminary extn. at room temp. Various details of app. and procedure are described.

30—RUBBER AND ALLIED SUBSTANCES

C. C. DAVIS

The chemical technology of rubber. SIDNEY M. HAGMAN. *Reprint from Svenska Gummis Ind. Memorial Number*, issued on the occasion of Disponent Henry Dunks' 60th anniversary, September 6, 1930.—A general discussion. H. C. DUUS

A modern service laboratory—the I. C. C.'s work for the rubber industry. ANON. *Chem. Trade J.* 52, 100-1 (1931).—A description is given of the new service lab. for the rubber industry that has been equipped by the Brit. Dyestuffs Corp., Ltd., at Hexagon House, Blackley, near Manchester. One section deals with the solution of technical problems, the second with the routine testing of rubber industry chemicals and colors and the third with research. W. H. BOYNTON

The world's rubber supplies. GEO. RAE. *Bull. Rubber Growers' Assoc.* 13, 59-84 (1921).—Though chiefly economic, the paper is of general interest, as it deals with planted areas, output, exports, imports, consumption and stocks in the producing and mfg. localities. C. C. DAVIS

Metallurgical methods used in producing rubber. PAUL D. V. MANNING. *Chem. Met. Eng.* 38, 131-2 (1931).—An illustrated description is given of new machinery for planting, cultivating and prep. guayule. C. C. DAVIS

Patents and custom standards of new machines for the manufacture of rubber articles. RUDOLF DITMAR. *Kolloid-Z.* 54, 237-9 (1931).—A review. A. FLEISCHER

The use of pigments in the rubber industry. FERD. A. RICHTER. *Farbe u. Lack* 1931, 135.—The primary use of pigments in rubber is not for the color but for the phys. properties imparted to the rubber. The choice of pigment depends upon the method of vulcanization as well as the use to which the product is put. G. G. SWARD

Reinforcing action of pigment mixtures on rubber compounds. D. J. BEAVER AND J. W. MACKAY. *Ind. Eng. Chem.* 23, 294-6 (1931).—The exptl. data which are presented show that mixts. of varying proportions of either channel gas black or a "soft" C black with whiting, lithopone or clay impart additive phys. properties to the vulcanizate. Mixts. of "soft" C black with ZnO also impart additive properties, whereas mixts. of channel black with ZnO result in poorer resistance to abrasion, in a higher modulus and in a higher tensile strength than those corresponding to the additive effects. These facts are explained by a chem. reaction between the basic ZnO and the acidic compds. in the rubber or on the C black. These results are applied to the formulation of a solid tire rubber mixt., which shows a greater resistance to abrasion and to "blow-out" when it contains a "soft" black than with a channel gas black. C. C. D.

American-grown rubber produced from guayule. GEO. H. CARNAHAN. *Chem. Met. Eng.* 38, 128-31 (1931).—An illustrated description of recent developments (cf. Spence, C. A. 24, 2914). C. C. DAVIS

Physical tests of sponge rubber. F. L. HAUSERALTER. *India Rubber World* 83, No. 5, 59-60 (1931).—An app. to measure the compressibility and the permanent set after compression of sponge rubber is described and illustrated. C. C. DAVIS

Soft rubber in chemical process equipment. H. E. FRITZ. *Chem. Markets* 28, 273, 275, 277 (1931).—An illustrated description is given of the application of rubber to ball mill linings, and machinery bearings, with special reference to the Vulcalock process. C. C. DAVIS

Chemical constitution of rubber. MARIO D. PINTO. *Rev. brasil. chim.* 2, 267-70 (1930).—P. discusses the work of the early investigators, and shows the evolution of the present formula (C_4H_6). JOHN M. LADING

Isoprene and rubber. XXII. Isorubber-nitrone. H. STAUDINGER AND H. JOSEPH. *Ber.* 63B, 2888-99 (1930); cf. C. A. 24, 4954.—Pummerer's view that the unit mol. of rubber is (C_4H_6), is supported by mol.-wt. detns. on rubber in menthol and especially on isorubber nitrone (I) in C_4H_6 and $PhNO$. These results, which have been confirmed as to their order of magnitude, are apparently not in agreement with Staudinger's view, based on model expts. with polystyrenes, the conversion of rubber into col-

laid sol rubber phosphonium salts, the prepn of polymer homologous series of poly-
prenes, the degradation of rubber, the relationship between viscosity and mol wt.,
etc., all of which lead to the conclusion that the primary colloid particles of rubber,
i. e., the particles in dil soln., are macromols consisting some 1000 isoprene mols. joined
in a chain (mol wt. around 68,000). A study of the action of PhNO on rubber solns
gave results somewhat different from those of Pummerer and Gündel. They believed
that the resulting decrease in viscosity is due to the disturbance of the rubber micelles
It has been found, however, that the viscosity of solns of macromol (eucolloid) poly-
styrene is greatly diminished by Cl_2 , Br_2 , strong oxidizing agents (N oxides, O_3), etc.,
as the result of an irreversible cleavage of the long sensitive macromols into shorter
fragments having a hemi-colloidal character. The viscosity changes (often considerable)
produced in highly viscous rubber solns by the addn of reagents are also irre-
versible, and therefore cannot be due to adsorption of the reagent on the colloid par-
ticles or to changes in their micellar structure. The macromols of rubber, because of
their double bonds, are much more sensitive than the said macromols of polystyrene,
and their cleavage into shorter fragments of hemi-colloidal character is a primary phase
of almost all transformations undergone by rubber. To det. the extent of such a cleav-
age in the reaction with PhNO , varying quantities (0.01-3.0 mols) of PhNO were added
to a 0.2 M soln of rubber (purified according to Pummerer) in C_6H_6 , and after 24 hrs
at 20° the time of flow from an Ostwald viscometer was noted. Parallel expts. were
made with NO_2 and Br_2 . Below are, resp., the no. of mols. of reagent used and the
viscosity of the resulting soln as compared with pure C_6H_6 taken as 1: PhNO 0.01,
3.2 3.0, 1.4 NO_2 0.01, 2.1; 0.03, 1.9, 0.1-1, product pptd out. Br_2 0.01, 3.6,
3.2 3.0, 1.4 NO_2 0.1, 3.7, 1.37. Rubber soln alone 3.6. If I, then, is not a deriv
of rubber itself but of oxidative degradation products, it should contain somewhat more
O than that calcd for pure isorubber nitron, and such is the case. In P and G's
published analyses only 2 show C values equal to the theoretical, the other 5 being
1-1.5% too low. S and J in about 20 analyses obtained values 1.5-2% too low al-
though they completely excluded air in the prepn of the I and carried out the filtration
and washing in a N atm. As to the mol wt. of the product, a series of detns. in 1%
 C_6H_6 soln gave values about twice as large as in PhNO , but the latter, which on the
av. were only slightly higher than P and G's, varied considerably. I is not a homo-
geneous substance but a mixt. of polymer homologous products which can be sep'd
by means of cold C_6H_6 into more easily and less easily sol. portions of approx. the same
compn. but with different mol wts. in PhNO , and forming CHCl_3 solns of different
viscosities. Furthermore, it was shown by isothermal distn. in Signer's app. (C. A. 24,
4433) that in equally concd solns of the 2 portions there are fewer mols. in the soln
of the less sol. portion. The reaction with PhNO does not permit of detg. whether the
mol. of rubber is made up of 100 or 1000 isoprene residues. XXIII. Cryoscopic measure-
ments on rubber solutions. H. STAUBINGER AND H. F. BONDY. *Ibid* 2500-5.—Reply
to Pummerer, Andriessen and Gündel (C. A. 24, 1102). Mol.-wt. detns. on rubber
in camphor cannot serve to clear up the structure of rubber, for, as already shown, the
rubber undergoes extensive degradation to hemi-colloidal cleavage products when heated
in molten camphor ($170-80^\circ$). The ratio of the time of flow of a C_6H_6 soln. of rubber
which has been dissolved in molten camphor to the time of flow of a soln. prep'd from
the same quantities of rubber and camphor without heating is about 1.15, the rubber
recovered from the former soln. has the properties of a deeply degraded hemi-colloidal
rubber, is smeary, has almost completely lost its viscosity, dissolves without swelling
and forms solns of low viscosity. The relative viscosities of a 0.2 M soln. of pure rubber
in tetralin are 14.89 and 2.60, resp., before and after the rubber has been heated with
camphor. The sep'n. of mixed phases observed by P. and Andriessen in attempts to
det. the mol. wt. of rubber in freezing C_6H_6 is easily explained by the fact that concd
rubber solns in C_6H_6 (1%) are not sol but gel solns., for rubber purified according to
P. the limiting concn., i. e., the transition from sol to gel soln., is about 0.5%, and it
is only in the dil. sol. solns. that mol.-wt. detns. can be made successfully. With the
hemi-colloidal degradation products, mol.-wt. detns. can be made by the 1 p. method
the limiting concn. of a degraded rubber of polymerization degree 100 is reached in an
approx. 5% soln. P and Gündel express the opinion that S's rubber was not com-
pletely N free, but S and B were unable to detect any N by heating with K , although
a control showed that 0.01% N can be detected in this way. C. A. R.

Swelling pressure of rubber. PAUL STAMBERGER. *Nature* 127, 274(1931); cf
C. A. 24, 4659.—The formula $P = K/V^2$ gives satisfactory results with the data of the
previous paper, where P is the swelling pressure, V is the vol. of solvent bound to a
unit wt. of jelly, and K is a const. whose meaning is expressed by substituting the above

equation in the max work term $dA = PdV$. The swelling pressure results from the attraction of solvent mols by the mols of the gel. ARTHUR FLEISCHER

The action of lead selenide in rubber mixtures. R. DITMAR AND K. H. PREUSSER. *Caoutchouc & gutta percha* 28, 1544-50 (1931).—Recent interest in Se in the rubber industry led to a study of its derivs as possible valuable compounding ingredients. The present paper continues the previous work on Se red (cf C. A. 24, 4063, D and Preusse C. A. 24, 4659). Zn selenosulfide (5% Se) and K_2SeSO_4 had no notable effect in rubber mixts. either on the rate of vulcanization or otherwise, so PbSe was next examd. A finely ground natural product (hardness 2.5-3 d 8.2-8.8) was used. In rubber-S mixts, PbSe has no influence on the rate of vulcanization nor any reinforcing effect. On the mixing mill PbSe has a strong plasticizing action which becomes marked with 4% PbSe and accordingly this property may be utilized in prep. adhesive cements from rubber. In the mixt. crepe 100, S 3 PbSe 2, accelerator 2, whitening 30, the PbSe acts as a very powerful activator with the majority of com. accelerators (the results with 16 are tabulated), so that no other activator is needed, and it also reinforces the vulcanizates particularly in increasing their resistance to abrasion. Its action, therefore, resembles that of Se. On increasing the PbSe in an accelerated mixt to 100 parts on the rubber, the reinforcement increased continuously. In cures with S_2Cl_2 in CS_2 or with S_2Cl_2 vapor, PbSe had no action at all, except that because of its softening action (*loc. cit.*) the cure had to be prolonged. C. C. DAVIS

Liquid rubber and carpets. WEBSTER NORRIS. *India Rubber World* 83, No. 6, 56 (1931).—An illustrated description is given of the application of rubber dispersions in a new type of carpet. C. C. DAVIS

The curing of sheet rubber. T. E. H. O'BRIEN. *Trop. Agr. (Ceylon)* 75, 280-92 (1930).—Smoking of sheet rubber prevents undesirable surface stickiness. Unsmoked sheet contg. *p*-nitrophenol is inferior in aging properties to a like sample without *p*-nitrophenol. Smoked sheet after vulcanization is superior in aging properties to unsmoked sheet when both contain *p*-nitrophenol. A sheet weighing 1.6 lb. and measuring 23 x 18 in. is of correct thickness for efficient drying, combined with good appearance, and is also the most suitable size for packing in a standard rubber chest. J. O. H.

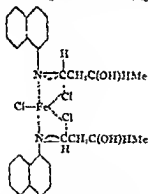
The consistency of the particles in balata latex. E. A. HAUSER. *Kautschuk* 7, 2-3 (1931). *Gummi Ztg* 45, 1030 (1931).—The expts. deal with balata latex from Peru, which was preserved with $AcOH$. For this reason the acid character of the latex may have brought about certain changes in the consistency of the particles, so that the conclusions drawn from the expts. do not necessarily hold good for the particles of fresh latex. The micromanipulator examn. was carried out with the aid of a new kind of dark field condenser, with the finest microneedle which was equipped with a new device for its control (cf C. A. 25, 1409). The examn. shows that there are 2 kinds of particles in such latex: (1) a uniformly tacky kind, which probably represents the resinous components of the latex, and (2) a smaller kind of particle which does not have a homogeneous structure. This 2nd type consists of a viscous inner part and a plastic outer part, which even has a certain degree of elasticity. By evapn. of the latex or by addn. of coagulating agents the consistency of the 1st type changes only slightly or not at all, whereas the 2nd kind of particle contracts to a homogeneous mass. The 2nd kind of particle probably therefore represents the actual balata hydrocarbon. Further expts. proved that during the evapn. only the particles of the 1st type fuse together, whereas the particles of the 2nd type remain distinctly sepd. from each other. This latter fact probably depends upon each hydrocarbon particle having an adsorbed skin of the non balata components dissolved in the serum. This observation is of particular interest, since it offers the possibility of obtaining pure balata by some convenient method, such as fractional pptn. or fractional sepn. from the latex, without having to resort to extrn. C. C. DAVIS

Gathering, treatment and properties of guttapercha. H. R. BRAAK. *Chem. Weekblad* 27, 567-71 (1930).—A review. A. L. HENNE

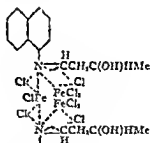
Age-resistant rubber. A survey of United States and foreign patents for chemical prevention of the deterioration of rubber. JOSEPH ROSSMAN. *India Rubber World* 83, No. 5, 65-8 (1931). C. C. DAVIS

The protective action of some antioxidants. II. The metal halide compounds of some protective agents against aging, with special reference to aldol- α -naphthylamine. F. KIRCHHOFF. *Kautschuk* 7, 7-12 (1931); cf C. A. 23, 3371.—It has already been shown that alc. solns. of heavy metal chlorides change the color of alc. solns. of certain antioxidants like aldol- α -naphthylamine (I), "Stabilite Alba" (II) and "Agerite powder" (III), both when viewed in visible and in ultra-violet light. The present paper, in which the behavior of I in particular was studied, attempts to explain this

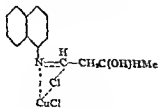
phenomenon, which is of importance not only from the antioxidant point of view, but because of the discoloration of vulcanizates contg such antioxidants. With AuCl_3 , CuCl_2 , FeCl_3 and PtCl_4 , I gives cherry red colors at low concns and intense blue colors at higher concns of the metal chlorides. The blue colors are, however, unstable and change to violet, then to cherry-red. With AlCl_3 , AsCl_3 , CoCl_2 , FeCl_2 , HgCl_2 , MnCl_2 , NiCl_2 , SbCl_3 , SnCl_2 , SnCl_4 , TiCl_3 , UO_2Cl_2 and ZnCl_2 , no colors are obtained. While AuCl_3 , CuCl_2 , FeCl_3 and PtCl_4 cause the disappearance of the fluorescence of I in ultra-violet light SbCl_3 is the only other chloride of those above which affects the fluorescence, and this gives a yellow white ppt. From the cherry red I solns with CuCl_2 and FeCl_3 , chocolate and coffee brown ppts with high Cu and Fe contents sep on long standing, and these remain dark even after long washing with I tOH. Their C_6H_6 solns are intense yellow brown. Analysis of the ppts showed that the Cu ppts, which vary from coffee brown to other yellow, are adsorption compds of the cherry red sol Cu I compds with solid I, and that the Fe ppts, which are chocolate brown, are adsorption compds of the sol cherry red Fe I compds with solid I. These reactions are not the same as that of aq I with aq FeCl_3 , ZnCl_2 , AgNO_3 and CrO_3 (cf "Meyer-Jacobson," II, 371), because water and dil alkalis cause cleavage of the alc Cu and Fe compds, with formation of blue violet flocculant ppts which contain no metal and which become yellow-brown in air. Since these color reactions are confined to chlorides of a higher state of oxidation the more loosely combined halogen plays the essential part. The fact that the reaction with CuCl_2 is in every case more rapid than with FeCl_3 may be utilized to distinguish analytically small quantities of these compds. A similar color reaction occurs with free halogen, e g, I_2 in CCl_4 , where a drop indigo appears immediately, changes to violet, then to cherry-red and finally to yellow-brown. These Br addn compds give no color reaction with FeCl_3 or CuCl_2 , so that in the color reaction the metal halide combines at the double bond between the N and CII. Since the cherry-red I FeCl_3 and I CuCl_2 compds are decompd by small quantities of HCl to yellow-brown products, only unstable compds are formed, and since the quantity of addn compd depends upon the concn of halide, the N does not always have the same valence, and the following formulas are assigned



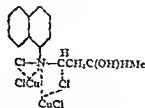
Cherry red soln (3 valence N)
from low FeCl_3 concn. Stable



Blue violet soln (4-valence N)
from high FeCl_3 soln Unstable



Cherry red soln (3-valence N)
Stable



Blue soln. (4-valence N)
Unstable

These resemble the hemin (IV) formula of Willstätter and Kuster. The I-metal halide solns absorb all visible light up to a bright band in the red and orange between 560 and 680 μ , and since HCl solns. of IV show a similar band, the absorption limits may be characteristic of the FeCl₃ and CuCl₂ groups. The ultra-violet fluorescence of all the I metal halide compds is very feeble compared with that of the original substance, and the colors vary greatly. The following fluorescent colors appeared in EtOH in ultra violet light: I, intense sky-blue, I + FeCl₃, very pale violet, I + CuCl₂, pale yellow green, I + CuCl₂ + HCl, pale green brown, I + Br, pale violet. II intense violet, II + CuCl₂, pale yellow green, II + CuCl₂, pale yellow green, II + CuCl₂ + HCl, pale green brown, II + Br, pale brown, IV + alc HCl, pale red brown. The source of the strong fluorescence is probably in the N—C≡N double bond, and it is lost by addn of the halide. The O transporting power of Fe in hemin occurs also in the I metal halide and II-metal halide compds, for I soln made pale rose by FeCl₃ becomes deep blood red when O is passed through it, whereas I soln alone is pale yellow. The O transporting power of the CuCl₂ compds is many times greater than that of the corresponding Fe compds and this agrees with the similar behavior of Fe and Cu compds in rubber. With 1% CuCl₂, O caused a turbidity with I and intense carmine with II, by the formation of oxidation products, and after prolonged passage of O, about 50% of the dissolved substances sepd as a chocolate brown ppt contg only a small percentage of combined Cu. Treatment of I soln contg FeCl₃ for many hrs gave an intense carmine color, with no insol oxidation products. Not only is the auto-oxidation of antioxidants greatly accelerated by their FeCl₃ complexes and especially by their CuCl₂ complexes, but it is inhibited by catalytic poisons. Thus with II CuCl₂ soln, a little KCN greatly retards the time of appearance and the intensity of the carmine which is formed upon treatment with O. III gives no immediate color with alc FeCl₃ or CuCl₂, but on long standing in air a brown color appeared, because of oxidation. III is probably more stable than I and II toward oxidizing agents and oxidation catalysts. In titrating III in ultra violet light, the quantity of FeCl₃ required to destroy the violet fluorescence is approx 2 mols of FeCl₃ per mol of III. II gives no immediate color with FeCl₃, but on long standing a carmine color appears, especially in light. II and CuCl₂ give an immediate yellow or yellow brown which changes soon to deep carmine. The carmine solns with II and FeCl₃ and CuCl₂ show similar spectral properties, and are not changed by HCl, in contrast to the colors with I complexes. In titrating in ultra violet light, only 0.5 mol of FeCl₃ or CuCl₂ is required to destroy the fluorescence of the II-FeCl₃ and II CuCl₂ complexes. Based on the similar absorption spectra of these metal halide complexes with the previous ones, they probably contain similar FeCl₃ and CuCl₂ complexes which cause the colors. The Fe complex

would be PhN C≡N, C≡N, NPh I Cl, where the stability toward HCl results from the formation of high valences. Since antioxidants are oxidized in time by autooxidation they protect an antioxidantizable substance like rubber only as long as they remain active. Since traces of sol heavy metal compds, especially of Fe, are invariably present in rubber compounding ingredients, rubber mixts which contain protective agents are discolored by the formation of dark colored metal complexes and their oxidation products. C. C. DAVIS

New light on vulcanization. The process as revealed by electrical tests. W. H. NUTTALL AND J. KIRKWOOD. *India Rubber J.* 80, 657-60 (1930).—A review and discussion, with particular reference to the work of Curtis, McPherson and Scott (C. A. 22, 330), Kitchen (C. A. 23, 2375) and Boggs and Blake (C. A. 24, 4125). C. C. D.

Tensile tests of vulcanized rubber at high speed. A. VAN ROSSEM AND H. B. BEVERDAM. *Rev gén caoutchouc* 7, No. 57, 27-34 (1930).—Most of the testing machines used for detg the elongation of rubber have a slow speed, and 60 cm. per min is somewhere near the speed usual in practice, regardless of the state of cure. This is in spite of the fact that undercured rubber, which has a high elongation when stretched at slow speed, is brittle when stretched very rapidly, and has an extremely low elongation. For this reason, it was of interest to make elongation tests at high speeds. No app was available for recording the stress-strain curve, and resort was had to the pendulum app of Charpy. The technic employed in adapting this app to the present expts is described and illustrated. Different accelerated rubber mixts at various states of cure were tested in this way. The energy at rupture was calcd, and the results were compared with tests on the Schopper dynamometer. The results show that (1) the max energy at rupture as a function of the time of vulcanization is much more sharply defined at the high speeds of the pendulum app than in the tests with the Schopper machine, and (2) the max strength with the pendulum app is at a

lower state of cure than with the Schopper machine. With vulcanizates contg a low percentage of S e g, the mixt. rubber 100, S 3, diphenylguanidine 1, ZnO 5, the pendulum app. gives a less sharply defined max., but nevertheless here too this max. is at a lower state of cure than that with the Schopper machine. The influence of loading with pigments and fillers was also studied. The magnitude of the max. depends upon the substance added, but the sharpness of this max. and its relation to the time of vulcanization are independent of the nature of the filler. Accelerated aging tests were made in the Gerr oven at 70° and the aged vulcanizates were then tested in the same way. After aging the time of cure/energy at rupture curve obtained from the Schopper machine was similar to that obtained from the pendulum app. before heating. The energy at rupture curve detd. by the pendulum app. is therefore of great importance in relation to the state of cure, because the time of vulcanization which corresponds to the max. energy at rupture by the pendulum app. is, contrary to the time of vulcanization corresponding to the max. tensile strength by the Schopper dynamometer, the time of vulcanization which corresponds to the best aging. The importance of tests at the high speeds and the shortcomings of the pendulum app. for this purpose make it desirable that an app. be developed which at a very rapid speed of stretching will: (1) det. the resistance to stress, the ultimate elongation and the energy at rupture (2) record graphically the relation between the stress and the elongation and (3) perform at different speeds. Also in *Kautschuk* 6, 224 9(1930). C. C. DAVIS

Hard spots in vulcanized rubber compounds. J. H. HOWEY, *Ind. Eng. Chem.* 23, 287-90(1931).—1 xpts. are described, the object of which was to ascertain the cause of local variations in the hardness of exptl. slabs of rubber mixts. of the C black tread type. It was found that areas of excessive hardness correspond to regions where the max. flow or displacement occurs when the rubber mixt. is pressed into the mold at the time of curing. A knowledge of this fact is of importance in testing the hardness of a vulcanizate by the Am. Soc. of Testing Materials standard of hardness, and for increasing the precision and reliability of this convenient test. It was also found by chem. analysis that these variations in hardness cannot be attributed to heterogeneous regions (non uniform compn.) in the rubber S-accelerator pigment mixt., and it is concluded that the variations must depend either upon altered conditions of dispersion in the regions of max. displacement or else upon a local grain effect set up by the flow. C. C. DAVIS

Use of vulcanized rubber waste as a fixing agent in bronze printing. Geo. RICE *Rubber Age* (N. Y.) 28, 611-2(1931).—An illustrated description. C. C. DAVIS

Monoxanthogens [vulcanization accelerators] (U. S. pat. 1,796,972) 10. Substitutes for sugars in coagulating latex. (Ger. pat. 518,195) 18. Dispersion of carbon black in water (Can. pat. 309,276) 18.

Purifying latex. K. D. P., LTD. Brit. 338,766, Dec. 20, 1923. Latex is subjected to heat treatment with simultaneous evapn. and then subjected to a usual purifying treatment, as by centrifuging to sep. serum, which is facilitated by the pretreatment.

Device for spraying rubber latex by an air jet. CHARLES E. MAYNARD (to Fisk Rubber Co.). U. S. 1,795,875, March 10. Structural features.

Rubber. K. D. P., LTD. (Rudolf Fumerner and Hans Kroepelin, inventors) Ger. 515,603, May 26, 1926. Addn. to 460,950. Pure rubber hydrocarbons are prepd. from rubber latex by treating with alkali in an electrolyzer.

Rubber compositions. DUNLOP RUBBER CO., LTD., E. A. MURPHY, A. NIVEN and D. F. TWISS. Brit. 338,975 Aug. 30, 1929. Crumb-like rubber compns. are prepd. by coagulating an aq. dispersion of rubber or the like, by pptn. *in situ* (in relatively large amounts on the dry rubber content) of compounding ingredients formed by the interaction of one or more water-sol. reagents normally having no coagulating effect, with other added water-sol. reagents, such as reagents which together form products such as $MgCO_3$, $CaCO_3$, $ZnCO_3$, Ca , Ba or Zn silicates, silicic acid or Zn , Cd or Pb sulfides. Soap, glue, gum acacia, etc., may be added to modify the properties of the product.

Use of amines as rubber coagulants. ALBERT J. GRACIA (to Goodyear Tire & Rubber Co.) U. S. 1,797,192, March 17. Latex is coagulated by use of amines such as ethylamine, diethylamine, propylamine, dipropylamine, tripropylamine, butylamine, amylamine, iso-amylamine, diisoamylamine or allylamine.

Preserving rubber. LORIN B. SEBRELL (to Goodyear Tire & Rubber Co.). U. S.

1,797,179, March 17. Amino-hexahydeneamine (suitably in the proportion of about 0.5%) is added to rubber compous before vulcanization

Rubber "antioxidant." Wm S CALCOTT and Wm A DOUGLASS (to E. I. du Pont de Nemours & Co) U S 1,796,980, March 17. Deterioration of rubber is retarded by the addn of a small proportion (suitably about 0.5-2.5%) of a transparent brittle solid obtained by the reaction of glucose on *m*-tolylenediamine

Trinaphthylamine. ALBERT M CLIFFORD (to Goodyear Tire & Rubber Co) U S 1,797,196, March 17 This compd is obtained by extrn with C_6H_6 from crude dinaphthylamines and is an antioxidant for rubber

"Micro-porous" rubber. H. BECKMANN Brit. 338,698, Oct 26, 1928 The process described in Brit 240,430 (C A 20, 2262) is modified by adding to the latex a substance capable of being coagulated to form a hydrophilic gel, such as 2% of blood albumin, and then adding a coagulating agent, such as HIOAc, which will cause coagulation of both the latex and the added substance Among the substances which may be added are silicic acid, stannic acid, $Al(OH)_3$, colloidal Fe_2O_3 , colloidal chromic oxide, colloidal tungstic and molybdic acids, manganese hydroxide, V pentoxide albuminoids, polypeptides and hemoglobin, and substances such as silicic acid may be formed in the latex by reaction Cf C A 24, 752

Apparatus for molding rubber articles such as tennis balls. HARRY WILLISIAW and SYDNEY N GOODHALL (to Dunlop Tire & Rubber Corp) U S 1,795,917, March 10 Various structural features are described

Uniting rubber and leather. K LIOKKE Brit 338,789, Jan 8, 1930 Materials such as shoe parts of rubber are roughened on a grinding wheel, coated with rubber soln, sprayed with particles of crepe rubber and further coated with rubber soln and then united by hammering with leather coated with rubber soln

Attaching rubber to metals. C. M. CARSON (to Goodyear Tire & Rubber Co) Brit. 339,421, Feb 23, 1929 A cement is used consisting of latex, hemoglobin a tanning agent such as formaldehyde, $K_2Cr_2O_7$, $K_4FeC_6N_8$ or $Al_2(SO_4)_3$, and vulcanizing ingredients and this cement is vulcanized after application

Compositions containing rubber and cement. N SWINDIN and NORDAC, LTD Brit 339,002, Sept 5, 1929 In making a material suitable for floors, road surfaces or acid resisting coating, raw rubber (heated or swelled with a solvent) is dispersed with water (preferably in the presence of a hard porous inert powder such as fine coke dust which acts as a trituring agent), suitably with addn of emulsifying agents such as soap, saponin, borax, NH_3 , and soda, and the material is then mixed with a cement (such as portland or oxychloride cement materials) which reacts with the water and causes setting of the mass Various details and use of different fillers, etc., are described

Films, threads, disks, etc. I. G. FARBERIN A-G Fr 37,485, Oct. 3, 1929 Addn to 676,658 (C A 24, 3135) Mixts of natural rubber with the polymerization products of butadiene hydrocarbons are treated with sulfuring agents until the high elasticity characteristic of rubber disappears

Isomerizing rubber. IMPERIAL CHEMICAL INDUSTRIES, LTD, and E B ROBINSON Brit 339,398, Sept 30, 1929 Rubber is liquefied by heat and then treated with isomerizing substances such as H_2SO_4 , sulfonic acids, sulfonyl chlorides, sulfuric esters or metallic halides, with or without phenolic compds, to produce a thermoplastic product Various details of procedure are given

Apparatus for curing tire flaps. ROBERT W SNYDER (to Goodyear Tire & Rubber Co.). U. S. 1,797,180, March 17. Structural features.

Rubber conversion product. HERMAN A BRUSON (to Goodyear Tire & Rubber Co). U. S. 1,797,188, March 17 See Can 299,963 (C. A. 24, 3135).

Synthetic rubber. I. G. FARBERIN A-G Brit 338,534, Aug 17, 1929 Polymerization of diolefins such as butadiene is effected in the presence of alkali or alk earth metals or their alloys in particles of uniform size, such as with Na balls of 1 mm diam in a rotating Fe autoclave at 40°.

Synthetic rubber. I. G. FARBERIN A-G. Brit. 339,135, Dec 20, 1929 Polymerization of diolefins such as butadiene with an alkali metal such as Na is effected in the presence of at least 25% of a low b-p aliphatic ether such as dimethyl, diethyl or methylethyl ether (these ethers serving both as diluents and as catalysts). Various details and examples are given

Synthetic rubber. I. G. FARBERIN A-G. Brit 339,243, Aug 30, 1929. Polymerization of diolefins such as butadiene is effected in the presence of alkali or alk earth metals or their alloys such as Na amalgam or of alkali or alk. earth metal compds with org radicals which are not capable of ionization (such as Na ethyl, K ethyl, Kd_1 phenylmethylmethane, Na triphenyl methane and Li alkyls) and the polymer-

ization agent is added gradually or in at least 3 successive portions during the polymerization. The polymerizing agent may be mixed with other substances such as polymers of diolefin rubber, paraffin, salt or waxes. Several examples are given.

Synthetic rubber. I. G. FARBENING A-G. Brit. 339,255, Aug. 30, 1929. Rubber like products suitable for various purposes similar to those for which rubber is usually employed are obtained by polymerizing a mixt. of a diolefin such as erythrene, isoprene, butadiene methyl or dimethyl butadiene, with styrene (suitably after emulsification of the initial materials together with Na stearate or other suitable emulsifying agent). The polymerization may be carried out at 40-60° for several days, and products of different properties may be produced by varying the proportions of the starting materials.

Synthetic rubber. I. G. FARBENING A-G. Fr. 37,374, Aug. 29, 1929. Addn. to 655,217 (C. A. 23, 4103). Synthetic rubber, etc., is intimately mixed with finely divided soot such as lampblack or C black before vulcanization.

Thiuram disulfides. GEORGE C. BAILEY (to Roessler & Hasslacher Chemical Co.) U. S. 1,796,977, March 17. Comps. suitable for use as rubber vulcanization accelerators and having the general formula $R_1CS-S-S-CS-R_2$ are made by treating a compd of the type $R_1CS-S-M$ with NaOCl, R_2 comprising an org. radical and M a metal radical. Tetramethylthiuram disulfide is made by treating Na dimethyldithiocarbamate with NaOCl in aq. soln.

Rubber vulcanization accelerators. RUBBER SERVICE LABORATORIES, INC. Brit. 339,332, Sept. 7, 1929. Accelerators are prepd. by reaction of an aldehyde amine such as butylidene isomylamine with a substituted unsatd. aldehyde such as α -ethyl β -propylacrolein. Various details of procedure for manuf. and use of the accelerator are described. Cf. C. A. 24, 266,752.

Apparatus for vulcanizing automobile tires. MATTHEW REID. U. S. 1,795,680, March 10. Structural features.

Vulcanizing rubber. INDUSTRIAL PROCESS CORP. Brit. 339,303, Nov. 26, 1928. See Fr. 681,031 (C. A. 24, 4126).

Vulcanizing rubber. HAROLD GRAY (to B. F. Goodrich Co.) U. S. 1,796,018, March 10. A product formed by reaction of an org. amine such as aniline or NH_3 , 1, and an aliphatic aldehyde such as butyric aldehyde 3 or more mol. proportions is used as an accelerator.

Rubber vulcanization. ADRIEN CAMBROU (to Roessler & Hasslacher Chemical Co.) U. S. 1,796,240, March 10. An accelerator is used comprising the product obtained by causing aniline and acetaldehyde, under essentially anhyd. conditions, to react in the presence of a small quantity of $ZnCl_2$. Cf. C. A. 24, 2640.

Vulcanized colored rubber. RUDOLPH KRECH (to I. G. Farbenind. A-G). U. S. 1,796,656, March 17. Before vulcanization, there is incorporated with rubber an insol. metal salt such as the Ba salt of an acid anthraquinone dye contg. at least one amino group, such as 1-amino-4-phenylaminoanthraquinone-2-sulfonic acid, and the material is then vulcanized.

CHEMICAL ABSTRACTS

Vol. 25.

JUNE 10, 1931

No. 11

1—APPARATUS AND PLANT EQUIPMENT

W. L. BADGER

Automatic laboratory siphon. CHARLES BUSHNICEVICH *Chemist-Analyst* 20, No. 3, 19(1931).—Glass caps, sealed to the outer sides of the ends of an equal-armed siphon tube, are recommended. W. T. H.

Simple technic in using a pycnometer. F. L. KISKADDEN *Chemist-Analyst* 20, No. 3, 16-7(1931).—Directions are given for using a narrow-neck, sp-gr bottle with viscous oils, tars and pitches. W. T. H.

A pipet bulb. R. W. KEITH. *Chemist-Analyst* 20, No. 3, 21(1931).—Rubber tubing, with wire to prevent collapse at the bend, can be used for making a suitable bulb. W. T. H.

A method for filling burets by vacuum. C. W. EDDY *J Chem Education* 8, 964-5(1931). E. H.

New insulator for the hot-water bottle. C. W. EDDY *Chemist Analyst* 20, No. 3, 20-1(1931).—Cut a strip of crude rubber of suitable size. Soak the edges in CCl_4 until they soften. Place around the neck of the bottle and work the edges together till the seam disappears. W. T. H.

Handy wash bottle. E. H. HERMAN *Chemist-Analyst* 20, No. 3, 21(1931).—The continuous flow wash bottle with Hunsen valve is described. A pressure bulb is used for blowing, and a glass head is placed in the rubber tubing which leads to the extra opening in the stopper. By pressing the rubber away from the head, air enters and the flow stops. W. T. H.

A thermometer for precision calorimetry. T. CARLTON SUTTON *J. Sci. Instruments* 8, 98-103(1931).—A Hg-in-glass thermometer for sensitive calorimetric work has been developed by (1) reducing to 60 mm the pressure of N in the space above the Hg column, and (2) eliminating parallax by means of a novel and simple type of reading lens. E. H.

Process for the investigation of air filters. W. SELL. *Z. Ver deut Ing* 75, 295-6(1931).—Previous methods for detg the efficiency of air filters have been gravimetric. A new colorimetric method for detg residual dust content is described in which a finely pulverized, water-sol. aniline dye is used as "colored dust." An accuracy of 0.5% is easily obtained. M. McMAHON

Accurate air separator for fine powders. PAUL S. ROLLER *Ind. Eng. Chem., Anal. Ed.* 3, 212-6(1931).—An app is described for sepg quantitatively a 1-kg charge of fine powder into a series of fractions beginning with 0-25 microns. Except for the finest fractions (below 5 microns), where attrition by the air current takes place in the case of soft powders, particle sizes sepd are homogeneous within the limits given by Stokes' law. The most important cause affecting the rate of sepn is that of rate of air flow. Under similar conditions rate of sepn is proportional to air flow. Initial rates of sepn of particle-size fractions have been made up to 135 g per hr at a flow of 140 l per min. Continuous sepn can be effected by the use of an offset separator tube with sep collection of oversize. A figure of the air separator and details of its operation are given. At a 30-micron particle size of a portland cement powder with an air flow of 500 l per min, the equiv. rate of feed was 54 kg per hr, while the rate of recovery was 0.92 kg. per hr. The efficiency of recovery, based on the max possible recovery in a run, was not in excess of 23%. ALICE W. ERPERSON

A simple procedure for fractional distillation of small quantities in a high vacuum. S. KONER *Biochem. Z.* 232, 274-7(1931).—A special app is described. S. MORGULIS

An apparatus for the evaporation of various materials in high vacuums. C. HANLEY CARTWRIGHT AND JOHN STROYA *Rev. Sci. Instruments* 2, 189-93(1931).—An app. for the deposition in high vacuum of metals and non-metals on various surfaces is described, in which material to be evapd is placed in a cooled filament and brought in contact with a heating coil in the evacuated app. The filament consisted of W or Pt wound

around a cylindrical dr. A crucible could be used if material to be evapd. reacted with W. or Pt. Al, Sn, In, Fe, Ca, Cr, Co, Cu, Au, Fe, Pb, Mg, Mn, Ni, Se, Ag, Te, Sn and Zn were deposited, also some alloys, quartz, fluorite, alkali halides and AgCl. on any surface not affected by high vacuum. M. McMIHAN

Glass extraction apparatus. B. L. MANGUNATH. *Half-Yearly J. Mysore Univ.* 4, 243-4 (1931)

GEORGE CALINGBERT

Extraction apparatus for solutions lighter than the solvent. ERISTO MAMMILLI. *Rev. chim. applicata* 21, 39-41 (1931).—A tube A with a siphon reaching almost to the bottom is placed in an ordinary Soxhlet extractor, and the soln. to be extrd. is placed in this tube, except for enough solvent just to cover the mouth of the siphon. The solvent evap. is condensed, and drops through A, which holds the soln. to be extrd. When the liquid reaches a certain level, it siphons out as in the ordinary app. When the vapor of the solvent is to be used, then a tube B brings the vapor up, discharging below the level of the soln. being extrd. Here also the condensed solvent siphons out automatically. A. W. COFFINER

A new absorption apparatus. ST. PRINER. *Chem. Ztg.* 55, 200 (1931).—The app. described was designed for studying the adsorbing and absorbing powers of activated C, etc., for solvent vapors. J. H. MOORE

The construction of Dewar flasks. T. E. FRYER, M. J. CONLEY and E. J. SNOW. *J. Am. Chem. Soc.* 53, 1355-6 (1931).—Details are given for using a 5-l round bottom pyrex flask for the outer jacket and a 3-l flask inside. These containers are recommended for the storage of liquid air. B. A. SOTLE

The expansion hydrometer. E. HEDERMANN. *Physik. Z.* 32, 223-5 (1931).—A new practical app. is described for the detn. of abs. values of the expansion coeff. of liquids. C. E. P. JEFFREYS

An automatic sampler and divider. H. GERICKE. *Chem. Fabrik* 1931, 159-60.—A detailed description is given of an app. for sampling or mixing liquids. It is driven by clock-work, will take as many as 24 samples, and the speed is adjustable to run for a day or a week. J. H. MOORE

Set of weights for the microchemical balance. WILHELM FELGENTRAGER. *Z. anal. Chem.* 83, 422-3 (1931).—The difficulties involved in getting satisfactory wts. which will be accurate to 0.001 mg. are pointed out. A description is given of a satisfactory set of weights for the microbalance and a method of calibrating the same to 0.001 mg. is described. W. T. H.

An apparatus for potentiometric volumetric analysis. PETER DICKEBUS. *Chem. Fabrik* 1931, 145-7, cf. *C. A.* 24, 1054.—Details are given of a simple and practical app. designed for the detn. of the elements usually found in steel alloys. J. H. M.

Apparent-specific-gravity apparatus. ARTHUR F. SUTER. *Eng. News-Record* 106, 522 (1931).—A brief illustrated description of a simple, inexpensive device for detg. the apparent sp. gr. of stone, slag, gravel, etc. R. E. THOMPSON

The efficiency of some common return condensers. F. MACH and R. HEDERMANN. *Chem. Fabrik* 1931, 157-9, 170-3.—Results are tabulated for expts. with 9 types of cooler when condensing Et₂O. The best 3 coolers in order of efficiency were Liebig coolers with 4 internal tubes, those with balls blown in the internal tube, and those with small pocket dams. Counter-current cooling did not increase the efficiency. J. H. M.

Magnetic probe in combustion tubes. P. A. THOMPSON. *Z. anorg. Chem.* 44, 278 (1931).—In Becco's recent paper (*C. A.* 25, 1453) a method for introducing the substance into a combustion tube is described, which is similar to that invented by T. (Microchemie 1, 6 (1924)) which is once more depicted. W. T. H.

Colloid milk. H. ARTHUR CHWALA. *Kolloid-Z.* 54, 358-64 (1931); cf. *C. A.* 25, 2722.—German patent review. A. FLEISCHER

And pumps. O. VOIT. *Metallwerke* 21, 555-6, 604-5 (1931).—Four types (centrifugal, plunger, diaphragm and those employing an auxiliary liquid) are described and compared as to construction, action, efficiency and use. Cast Fe contg. 15-18% Si is commonly the fabricating material, although bronze and hard Pb are not attacked by a no. of acids. The stuffing box is an important part of any pump. B. A. SOTLE

Automatic vacuum pressure regulator. HOWARD L. COLL. *Philippine J. Sci.* 44, 411-3 (1931).—A modified Miller and McKimsey app. (*C. A.* 22, 2294) to be used on a 110-v. d. c. or on a battery is described. The Hg of the manometer touches a Pt-wire contact, thus closing a circuit and bending down the movable arm of an electromagnet. This arm holds a rubber stopper with a hypodermic needle No. 22 inserted in it. As air enters the system through the needle, the current is interrupted by the falling of the Hg in the manometer. This device maintains a vacuum within ± 0.5 mm. at 20 mm. pressure. ALFRED BRIDGE

Design, construction and operation of a constant-humidity room. P. H. PRIOR *Proc. Tech. Sect. Paper Makers' Assoc. Gt Brit & Ireland* 9, 152-74(1929) —The points to be considered in the design of such a room are discussed, and details are given for the construction of a room 12 ft. by 15 ft. by 7 ft. 10 in. Provision is made for complete change of air 4 times per hr. or for recirculation of the air over the controls when the room is only used for conditioning samples. Humidity and temp control are effected by diverting part of the air over the correcting devices (humidifier, refrigerator and heater) by means of elec. fans controlled by a hair hygrometer and a bimetal strip. The controls normally operate at about 5-min. intervals and to within 1% relative humidity and 1° F. B. C. A.

Low-temperature test rooms an aid to research. HUGH J. KRAMPE *Refrigerating Eng.* 21, 331-3(1931) —Recommendations are given as to methods and materials for use in the construction of low temp test rooms. The uses of such low temp test rooms in the automotive, oil and food industries, as well as in academic research institutions, are described. A. H. JOHNSON

Oblique-light illuminator for photography by reflected ultra-violet and by the light of fluorescence. G. KÖGEL *Phot. Korr.* 66, 282-3(1930) —In the adaptation of the Busch oblique-light illuminator to ultra violet photography, K. suggests the use of Cr-plated metal mirrors (described and illustrated) instead of the usual, but relatively expensive, optical system of quartz. The special suitability of reflected ultra-violet radiation for photographing black lead pencil marks is mentioned. E. R. BULLOCK

The calculation of the thermal characteristics of regenerators. K. RUDOLPH *J. Inst. Fuel* 4, 160-73(1931) —The process of heat transfer in an exptl. regenerator was studied. Measurements of temp and resistance to flow were taken on different types of checkers and under varying conditions of operation. Practical simplified methods of calcn are derived and applied to the exptl. data. D. A. REYNOLDS

Glass blow-torch for blow-pipe analysis. C. H. ALMFELT AND C. C. BEEBE *J. Chem. Education* 8, 968(1931) E. H.

Monochromatic sodium light in polarimetry and refractometry. VL. STANĚK AND K. ŠANDERA *Listy Cukrovar.* 49, 311-5(1931) —The elec. Na lamp of the firm Osram is described in its simplest form, and its use in polarimetry is discussed. The light may be purified with filters or by prisms. The simple monochromator (Schmidt and Haensch) is attached to the polarimeter and works well with the above Na lamp. For prism polarimeters the filtration of light is not necessary. FRANK MARESH

Note on Julius suspensions. M. J. BREVOORT. *Bur. Mines, Rept. of Investigations* 3086, 2 pp (1931). —An improved galvanometer suspension is described. ALDEN H. EMERY

What is the basis for retiring equipment? R. L. CORSON. *Chem. Mfg. Eng.* 38, 202-3(1931). —The economics of replacement of equipment is discussed. E. H.

Machinery to make solid carbon dioxide. TERRY MITCHELL. *Ind. Eng. Chem.* 23, 523-5(1931) —The app. used in the manuf. of solid CO₂ is briefly reviewed. Illustrations and a flow sheet are included. R. D. BUMBACHER

A bath thermostat. THOS. C. HERNDON *J. Chem. Education* 8, 969-70(1931). E. H.

A thermostat regulator. ELLIOT Q. ADAMS. *Rev. Sci. Instruments* 2, 187-8(1931) —A W-Hg contact m. H. permanently enclosed in glass was used as a thermostat regulator. A Hg tight W plunger in 2 mm. capillary served as adjuster, one complete turn of the micrometer bead advancing the plunger approx. 1 mm., equiv. to 0.08°. The regulator was used for keeping a 50-l. H₂O thermostat const. to $\pm 0.002^\circ$ for 1 hr. With an auxiliary heater in a heating-coil circuit, variation in temp. did not exceed $\pm 0.001^\circ$ for 3 hrs. M. McMAHON

Permanent marking of glass and porcelain vessels (ISBELL) 19. A microscopic method [hypermicroscope] (SYNGE) 2. Theory of optical pyrometers and a description of several new improvements (HELD) 2.

TRINKS, W.: *Industrieföfen*, Band II. Bau und Betrieb. Berlin: V. D. I. Buchhandlung 398 pp. Linen, M. 20 Cf. C. A. 23, 2834.

Laboratory extraction apparatus (Soxhlet type). ARNOLD FLIEDNER. *Ger.* 519,371, May 26, 1929.

Thermometer. ILMENAUER GLAS-INSTRUMENTEN-FABRIK ALBERT ZUCKSCHWERTDT. *Ger.* 521,228, Jan 17, 1929. The scale-tube and the mercury (or other liquid) reservoir are connected by a flexible metal capillary tube.

- Dropping bottle. C F BOEHRINGER & SÖHNE G M B H Ger 519,476, June 15, 1929
- Filters. AUGUST PAHL Ger 469,227, May 21, 1925 See U. S. 1,780,381 (C. A. 25, 440)
- Filters. AUGUSTE G LEMAIRE Fr 696,403, Nov. 7, 1929 Arrangement for closing the filter pockets is described
- Filters. MOTOR IMPROVEMENTS INC Fr 697,001, June 3, 1930 A metallic filter made from a spiral ribbon of metal is described
- Air or gas filters. HENRI HAEGLER Fr 696,791, Sept. 17, 1929
- Cylindrical filter for air or gases HANS WITTEMEIER Fr 695,427, May 10, 1930
- Metallic filters for air or gases SAMUEL A NIESTLÉ Fr 695,445, May 12, 1930
- Gas filter. JULIUS A WILISCH Ger 490,245 April 16 1921
- Gas filter. R LAHAUSSOIS Brit 339,913 March 30, 1929 A glass plate or the like is coated with an adhesive and hairs are taken from a support and caused to adhere by one end to the prepd surface to form a filter plate for gases Cf C. A. 25, 1414
- Apparatus for separating suspended matter from gases HORACE WAXING (to Associated Lead Manufacturers, Ltd) U. S. 1,797,812, March 24 Structural features
- Device for separating liquids from gases. HARLAN W HOW (to Struthers Wells-Titusville Corp) U. S. 1,797,232, March 24 Structural features of a device suitable for sepp entrained liquid from vapors from evap app
- Apparatus for filtering liquids. ERNEST LAGGER Swiss 143,378, Jan. 20, 1930 The filter beds comprise layers of wood charcoal and gravel
- Sand filters. SOC H CHABAL ET CIE Fr 696,366, May 31, 1930 A method is described of cleansing by air blown through
- Drum filters. I G FAHRBUND A-G Fr 695,818, May 19, 1930 An endless band with tensioning means is wound in a helix round the filtering and heating or drying drum to remove solid substances retained on the filtering surface
- Continuously operating pressure-cell drum filter. FLORENZ FLEMING Ger 521,030, Dec. 15, 1928
- Pressure filter and extrusion apparatus. A R JANN Brit 339,669, Sept. 4, 1929 Structural features
- Water-film apparatus for removing dust from gases. LUDWIG DOLF Ger. 519,017, April 27, 1928
- Centrifugal separators for mixtures of three liquids. AKTIEBOLAGET SEPARATOR-NOBEL Fr 695,725, May 16, 1930
- Means for distributing blown air under filtering layers. SOC. H. CHABAL & CIE Fr 696,614, June 3, 1930
- Roller driers. KALI-CHEMIE A-G and WILHELM SASSER Ger 516,757, July 14, 1928
- Band driers for crystals obtained from solutions. ADLER & HENTZEN, MASCHINENFABRIK Ger 516,658, Jan. 4, 1929
- Tube system for pneumatic drier. PAUL ROSIN Ger 520,056, Jan. 25, 1929
- Drum and plate drier. MITTELDEUTSCHER STAHLWERKE A-G Ger 517,718, Oct. 22, 1929
- Drying drum. MASCHINENFABRIK IMPERIAL G M B H Ger 521,100, Sept. 24, 1929
- Drying drum with internal mixers. DEUTSCHE BARCOCK- & WILCOX-DAMPFERSEL-WERKE A-G Ger 517,716, Jan. 10, 1930
- Vacuum drying apparatus. JULIUS PINTSCH A-G (Walter Neuhold, inventor) Ger 517,719, Oct. 24, 1929 Details are given of a stationary surface condenser.
- Vacuum drying apparatus. JULIUS PINTSCH A-G (Walter Neuhold and Hans Groth, inventors) Ger 520,285, Oct. 24, 1929
- Drying kiln. RUDOLF WIEDEMANN Ger 517,908, Dec. 6, 1927.
- Rotary kiln or drying apparatus. F A JOHNSON Brit 339,695, Sept. 4, 1929 Structural features
- Centrifuge for drying vapors. WM. ALEXANDER Ger 516,553, April 30, 1927.
- Apparatus for drying air and other gases. ALEC E SHERMAN and ISIDORE MAYNARD Ger. 520,055, Aug. 14, 1929 Brit. 314,692 (C. A. 24, 1254)
- Drying plant comprising chambers and hot air currents. U FORTNITZ & Co. A-G Ger 517,715, Nov. 1, 1927.
- Automatic regulating means for evaporators. PRANTISEK TUREK Fr. 697,238, June 11, 1930

- 517 745 June 12, 1928 The hot gases are led partly under the base of the muffle and partly around its sides
- Coal-dust furnace CARL HUPFSCHMIDT Ger 511,093, Aug 11, 1926
- Gas furnace KURT RUMMEL and ALFRED SCHACK Ger 517,818, Nov. 8, 1928 Details of a metal recuperator for preheating the fuel to a high temp are given
- Furnace for the activation of carbonaceous materials by gases. SOC. DE RECHERCHES ET D'EXPLOITATIONS PÉTROLIÈRES. Fr 696,253, Sept. 9, 1929
- Heating regenerative furnaces. VEREINIGTE STAHLWERKE A.-G. Ger. 519,124, Sept 18, 1927 Coal dust is blown by means of a combustible gas into the hot gases passing to the furnace, the arrangement being such that the coal is partly or wholly gasified before it reaches the hearth
- Air heater for furnaces, etc. MAX and ERNST HARTMANN. Ger. 519,471, Aug 24, 1929
- Supplementary air-draft regulator for industrial furnaces. SCHMIDT-ARSOLO G M B H FÜR FEUERZUGREGELUNG Ger. 520,314, Mar 27, 1929
- Traveling-grate apparatus for drying, destructive distillation, etc. FRANZ SCHENK. Ger 520,075, Oct 31, 1928
- Device for preventing escape of hot gases through faulty brickwork of furnaces by suction. ERICH ROUCKA Ger 517,746, Feb 15, 1925
- Hollow water- or air-cooled beams for the fire chamber of furnaces. GUSTAV KARRENBERG and HERMANN RÖMER Ger 517,903, Sept. 30, 1928
- Burners for powdered fuel. GEORGE E. K. BLYTHE, Fr. 690,713, June 6, 1930
- Coal-dust burner. ADOLF BEULMANN Ger 519,426, Feb 11, 1927
- Burner for pulverulent or gaseous fuel. ALLGEMEINE ELEKTIZITÄTS-GES. (Friedrich Munzinger, inventor) Ger 519,022, Jan 14, 1928
- Gas burner. OSKAR HOPPE Ger 519,425, Aug. 26, 1927. Means are described for regulating the admixt. of the gas with compressed air
- Heavy-oil burner. DECIO PIZZI. Ger. 516,507, Nov. 6, 1929 Details of construction are given
- Heat exchanger. HANS SIMON Ger 519,054, Aug 19, 1925 See Austrian 113,190 (C. A. 23, 4109)
- Hest exchanger, especially for steam generator. SCHMIDT'SCHE HEISSDAMPF-GES. M. B. H. Ger 519,053, Sept. 23, 1928
- Tubular hest-exchange apparatus. R. MORTON & Co., LTD., and P. ROBINSON. Brit. 339,800-70, Feb 7, 1930 Structural features
- Apparatus for heating by means of gaseous or liquid fuel with recirculation of the active gases and elimination of the inert gases. PAUL G. LINKER. Fr 690,346, May 31, 1930
- Heat-transmission tubes for chemical apparatus. ROYAL BAKING POWDER CO. Ger 519,047, May 16, 1928. Brit. 311,092 (C. A. 24, 534)
- Control system for regulating the supply of fuel such as gas to boiler furnaces in accord with pressure variations. THOMAS A. PEEBLES (to John M. Hopwood). U. S. 1,797,586, March 24. Structural features.
- Prewarming and cooling apparatus for annealing pots. AET.-GES. BROWN, BOVERI & Cie Ger 519,330, May 20, 1930.
- Electric heater for liquids. EGERTON R. CASE. U. S. 1,797,520, March 24. Structural features.
- Electric water heater. A. LOW & SONS, LTD., and D. W. LOW. Brit. 339,734, Oct. 24, 1929 Structural features
- Electric water heater. THOMAS B. ALLARDICE U. S. 1,797,749, March 24.
- Acid proofing. PAUL ASKEVASY Ger. 516,745, Dec. 6, 1929 The rotating parts of centrifuges, etc., are rendered acid proof by coating with hard Pb contg. Hg
- Apparatus for evaporating brine, etc. PAUL H. MÜLLER. Ger 519,123, Aug 13, 1925
- Apparatus for making foam. GEORGE M. THOMSON Ger 521,183, Feb. 8, 1928 See Brit. 299,242 (C. A. 23, 3132).
- Moisture-condensing device. CHARLES W. ARMERUST (to International Life Saving Water Making Cup Corp) U. S. 1,797,894, March 24 Various details are described of a device "for condensing the moisture from exhaled breath," having for its stated object "the provision of a life-saving cup which will collect the moisture from the breath and condense it as pure water for and during emergency, as in shipwreck, or on the sea, or in desert places, without fresh water"
- Cooling-tower. K. W. BRANCZIK. Brit. 340,127, Dec 16, 1929
- Precision casting machine. FRIEDRICH J. HAAS. Ger 521,387, Feb 13, 1926

Die-casting machine for artificial material. INTERNATIONAL METALL A-G. Ger 517,510, Mar 25, 1928.

Device for removing the cores of castings. SIEMENS-SCHUCKERTWERKE A-G. Ger. 517,781, Dec. 15, 1928.

Apparatus for electroosmotic purification of liquids such as water. SIEMENS-ELEKTRO OSMOSE GES. Brit. 339,673, Sept. 7, 1929. Structural features and various details of operation are described.

Apparatus for evaporating corrosive solutions such as caustic soda. R. M. WINTER and IMPERIAL CHEMICAL INDUSTRIES, LTD. Brit. 339,657, June 14, 1929. The liquid to be evapd. is passed as a film over a surface heated directly or indirectly by radiation from one or more incandescent plates arranged parallel to the surface. The plates may be formed as described in Brit. 25,803 of 1909 or consist of a metal plate supporting an aluminum plate carrying granules of magnesia or the like. Various details of construction are described.

Lubricated rotary plug valve for fluids destructive to the lubricant. S. J. NORDSTROM (to Merco Nordstrom Valve Co.). Brit. 339,889, Feb. 26, 1929. Structural features.

Apparatus for grading tea or other materials by the action of air currents. W. G. FIRMAN. Brit. 339,758, Nov. 14, 1929. Various structural details are described.

Heated vacuum container with stirring device for mixing materials such as graphite and bearing metals. ERNST MEIER (to Braunschweiger Huttenwerk G. m. b. H.). U. S. 1,797,992, March 24. Structural features.

Mixing and emulsifying apparatus of the cylinder and plunger type. W. J. DAVY. Brit. 339,751, Nov. 8, 1929. Structural features.

Apparatus for testing the tensile strength of flexible materials. ALFRED SCHOPPER (to the firm of Louis Schopper). U. S. 1,797,734, March 24. Structural details of a gas inflation testing app. are described.

Apparatus and mode of operation for classifying sand or other powdered materials by aqueous suspension and centrifugal and gravity separation. L. ANDREWS. Brit. 340,027, Sept. 19, 1929.

Surface condenser. INGERSOLL-RAND Co. Ger. 520,462, Nov. 17, 1928. See Brit. 302,347 (C. A. 23, 4109).

Containers for storing gases or liquids under pressure. ALEXANDER BELDIANO. Ger. 517,917, June 27, 1928.

Carbide agitator for acetylene generators. HERBERT G. IRWIN. U. S. 1,797,264, March 24. Structural features.

Acetylene generator. CARL KEEL. Fr. 695,278, May 7, 1930.

Acetylene generator. I. G. FARBENIND A-G. Fr. 696,359, May 31, 1930.

Thermoelectro-responsive device suitable for control of switches, etc. HERBERT E. RUPP (to Ohio Brass Co.). U. S. 1,797,369, March 24. Structural and elec. features. U. S. 1,797,370 describes a thermostatic elec. circuit control device.

Bimetallic thermostatic electric switch or cut-out. FELTEN & GUILLEAUME A-G. Brit. 339,705, Sept. 6, 1929. Structural details.

Thermostatic device for control of electric circuits. ADOLPH A. THOMAS. U. S. 1,797,886, March 24. Structural features.

Thermostatic control device for gas fed to burners. OTTO FOX and WILLET E. STARK (to Bryant Heater and Mfg. Co.). U. S. 1,797,571, March 24. Structural features.

2—GENERAL AND PHYSICAL CHEMISTRY

FREDERICK L. BROWNE

Harley James Morrison. A. S. RICHARDSON. *Ind. Eng. Chem.* 23, 594(1931).—Biography with portrait. G. C.

Karl Seubert's eightieth birthday, April 6, 1931. WILHELM BILTZ. *Z. anorg. Chem.* 44, 269(1931).—Biographical note with portrait. E. H.

The investigations of Frant. Wald. ANTONIN KŘÍŽ. *Chem. Listy* 25, 5-10(1931); cf. C. A. 25, 1714.—A review. FRANK MARESH

In the services of chemistry. ZD. RAUŠAR. *Chem. Listy* 24, 437-43(1930).—An article filled with reminiscences covering the advances made in chemistry during the years 1880-1930. FRANK MARESH

The progress of physicochemical research in India. S. S. BHATNAGAR. *Proc. 15th Indian Sci. Congr.* 1928, 95-114.—An address. E. J. C.

A Chinese forerunner of the Bunsen burner. FIMMERICH CZAKÓ *Z. anorg. Chem.* 44, 22 (1911) — A device used by the Chinese for burning natural gas, reported by the missionary Imbert in 1828, is described. E. H.

The new chemical laboratories of Picatinny Arsenal. J. B. ROSE. *Ind. Eng. Chem. Anal. Ed.* 9, No. 8, 123 (1931) E. J. C.

The "Deutsche Museum" and its sections of chemistry and pharmacy. CARLOS DANTAS ARAUJO *Bol. Assoc. Brasil. Pharm.* 11, 25-33 (1930) — Descriptive. C. A. ABLEDO

Romance of science in bygone London. H. G. WATLING. *Science Progress* 25, 47-50 (1931) — A paper of considerable interest to students of the history of chemistry and the other phys. sciences. JOSEPH S. HEPBURN

Chemistry in the intermediate schools. JOSEF KOE. *Chem. Listy* 24, 448-50 (1930) — The defects of education are discussed and reforms are proposed. FRANK MARESH

An industrial chemistry text of 1830. CHARLES W. STILLWELL. *J. Chem. Education* 8, 896-901 (1931) — A review. E. H.

Some aspects of the unit method of teaching chemistry. RUSSELL S. HOWARD *J. Chem. Education* 8, 910-8 (1931) E. H.

Inspection trips. W. T. READ. *J. Chem. Education* 8, 919-28 (1931) E. H.

Some instructional devices in elementary chemistry. HORACE G. DENING *J. Chem. Education* 8, 902-9 (1931) E. H.

A study of the type and number of balances essential for high-school use. RALPH B. DENBAR AND EMORY D. FISHER. *J. Chem. Education* 8, 936-8 (1931) E. H.

Equation testing in high schools. LEONARD J. FLIEDNER. *J. Chem. Education* 8, 939 (1931) E. H.

A student experiment on the olefin hydrocarbons: The interconversion of ethylidimethylcarbinol and amylene; the polymerization of amylene. W. M. LEE AND E. C. WAGNER. *J. Chem. Education* 8, 941-5 (1931) E. H.

Hydrocyanic acid from methylamine as a lecture experiment. HERMANN EMOR AND THOR HORNEMANN. *Z. anorg. Chem.* 44, 278 (1931) E. H.

A reaction-velocity experiment. JACK P. MONTGOMERY. *J. Chem. Education* 8, 940 (1931) — For class-room demonstration, the course of the acid hydrolysis of sucrose is followed by the change in sp. gr. of the soln., measured by a Westphal balance. E. H.

The specific gravity and the vapor pressure of concentrated sea water from 0° to 175°. K. HIGASHI, K. NAKAMURA AND R. HARA. *J. Soc. Chem. Ind. Japan* 34, Suppl. binding 72 (1931) — The sp. gr. and the vapor pressure of sea water at concns. from normal to satd. NaCl and at temps. from 0° to 175° were detd. V. F. H.

Nomography. OTTO LIESCHKE. *Chem. Fabrik* 1931, 122 — The chart shows change in % with changing wt. during evapn. and drying. J. H. MOORE

Addendum to the Utrecht radiation and temperature measurements on the measurement of a black body by means of the gold point. L. S. ORNSTEIN, D. VERMETLEN AND J. WOUTA. *Proc. Acad. Sci. Amsterdam* 33, 985-9 (1930) — Temps. of W lamp filament were measured with a calibrated thermopile and an optical pyrometer, and the space distribution and emission coeff. calcd. To connect these measurements with the international scale a black body of Ni was inserted in a tube furnace. Just in front of this was placed a Pt-PtRh thermoelement clamped together with a small piece of very pure Au. The temp. was measured simultaneously with the thermoelement and a pyrometer, the interruption of the thermal current giving the m. p. of the Au. GREGG M. EVANS

Theory of optical pyrometers and a description of several new improvements. S. HELD. *Chaleur et Ind.* 11, 473-80 (1930). cf. C. A. 25, 1414 — Standardization, the action of selective filters, the transmission factor with a selective filter, an interpolation formula for selective filters, conditions for the use of optical pyrometers, calcn. of the temp. of a black body as compared with a non-black body, a practical formula for measuring black body temp., measurement of the temp. of a non-black body with selective screens and the possibilities of error, are treated mathematically and graphically. S. L. B. ETHESTON

Radiation pyrometry. E. L. FIDGR. *Chem. Eng. Mining Rev.* 23, 189-94 (1931). — The application and use of radiation pyrometers in industry and lab. research are discussed. F. D. ROSSINI

Methods of measurement with the thermionic tube. ERNESTO DENTINA. *Atti III congresso naz. chim. pura applicata* 1930, 402-13 — The application of the thermionic vacuum tube to problems of phys. chemistry and electrochemistry is described.

to the free energy change. A correlation is claimed between the order of increasing free energy change and the order of increasing abundance of the element formed.

The new element rhenium. WILHELM FRIT. *Ind Eng Chem., News Ed.* 9, No. 8, 122(1931).—Descriptive. J. B. AUSTIN
E. J. C.

Atomic weight of di-manganese (rhenium). J. G. F. DRUCE. *Science Progress* 25, 499-500(1931), cf. *C. A.* 25, 2031.—A review of detns of this at. wt by various investigators. The value 186.31 ± 0.02 is accepted as the most accurate. J. S. H.

Investigations for the concentration and isolation of the element having atomic number 61. L. ROLLA. *Ann. Acad. Lincei* 12, 270-2(1930).—A review of work done to date in attempting to isolate element 61 from a large sample of com. didymium, contg. considerable Ce. A. W. CONTIERI

Statistic evaluation of the energy of interaction of a coulomb and a molecule. L. GOLDSTEIN. *Compt. rend.* 191, 606-6(1930).—Math. GREGG M. EVANS

Some aspects of the problem of molecular structure. JAMES KENNER. *Rept. Australian Assoc. Advancement Sci.* 18, 132-55(1930)(published in *Sydney Univ. Reports Ser. III*, 1, Nos. 45-61).—A presidential address before the Australian Assoc. Adv. Sci. 18, Sect. B(1930). A general review. C. E. P. JEFFREYS

The physical identity of enantiomers. ALAN N. CAMPBELL AND FREDERICK C. GARROW. *Trans. Faraday Soc.* 26, 500-4(1930), cf. *C. A.* 24, 1000.—Extremely pure *d*- and *l*-mandelic acids are shown to differ slightly but distinctly in m. p., rotatory power, rotatory dispersion and soly. The hypothesis is proposed that enantiomers are not physically identical. A. P. SACHS

Symmetric apherical oscillator and the rotational motion of homeopolar molecules in crystals. T. E. STERN. *Proc. Roy. Soc. (London)* A130, 551-7(1931).—A treatment of a homeopolar mol. of type X_2 free to rotate in a field of force of axial symmetry is developed. The method is more rigorous than that of Pauling (*C. A.* 24, 5559) and more extensive. It is concluded that rotation occurs if T is greater than $2V_0/k$ and that oscillation occurs if T is less than $2V_0/k$, wherein T is the temp., k Boltzmann's const. and V_0 the potential function defined by P . Thus two cases are distinguished, one of rotation and one of oscillation. WILLIAM E. VAUGHAN

Studies on coordination. III. The energy of coordination. F. J. GARRICK. *Phil. Mag.* 11, 741-4(1931), cf. *C. A.* 24, 2024, 5578.—The energy of coordination is the min. potential energy (ϕ) or the max. energy evolved on formation of complex ion. Values of ϕ calcd. for Zn, Fe, and Mn as $MCl_2 \cdot 6NH_3$, by the Born theory of lattice energy agree with those calcd. by the previous method. Data for calcs. on hydrates are lacking. ARTHUR FLEISCHER

The deformability of ions. R. BRDICKA. *Chem. Listy* 24, 374-9, 398-403(1930).—The Fajan's theory is used to explain many phys. phenomena and chem. properties, the theory remains on a qual. basis but embodies a phys. concept of atoms and ions. The theory cannot be used to explain the color of salts as an absorption of visible radiations due to deformed anions, for the electronic system of cations also absorb in this range. Adsorption may be associated with the deformability of ions. The adsorption depends upon the potential at the interface of two phases. The potential originates in the Helmholtz double layer and depends upon the distance between the opposite charges in the layer. The more deformable ions permit a close approach of the electrons in a shell and force the pos. layer to approach the neg. layer. This increases the adsorption potential and consequently the adsorption. FRANK MARSH

Surface tension in a magnetic field. HERMANN AUER. *Z. Physik* 66, 224-8(1930).—No effect of a magnetic field of 20,000 gaussess could be detected within the accuracy of the expts. (10^{-2} - 10^{-4}). GREGG M. EVANS

Concentrated space charge (polarization) in calcite. A. JOFFE, D. ROYANSKIY AND K. SINTELNIKOV. *Z. Physik* 66, 143-71(1930).—The variation of space charge in the thin cathode layer was measured and expressed as the capacity of an equiv. condenser. The field at the cathode is so strong that a drop of Hg is flattened when used as an electrode. Measurements were made of (1) current with const. p. d., (2) potential with const. current and (3) residual current always reached as a min. (in reality a leakage current of the condenser). The capacity was, in general, higher parallel to the Z -axis than at right angles to it, but was unchanged with time. The cond. increased logarithmically with temp. The behavior of the different crystals and electrodes was individual. The space charge leaked away, even when no imposed current was passing, but took months for complete leakage at room temp. The concn. of space charge in calcite crystals in these cathode layers is attributed to impurities in the crystal. GREGG M. EVANS

Electrical conductivity and high voltage polarization of saltpeter crystals. BORIS HOCHBERG AND VALENTINE JOFFE. *Z. Physik* 66, 172-91(1930), cf preceding abstr.—The NaNO_3 crystal was made from a melt. The ratio of polarization e. m. f. to that imposed decreased with increase in temp and imposed e. m. f. The space charge is concentrated in thin layers at both electrodes and is attributed to impurities, since the polarization decreases with increased purity of the crystal. GRIGG M. BYAN.

Thermomagnetic study of some anhydrous compounds of cobalt and nickel. B. CARRERA AND A. DUTERIER. *Anales soc. esp. fis. quim.* 29, 5-14(1931).—The thermic variation of magnetic susceptibility of some Co and Ni compounds was determined by the Faraday method of electrodynamic compensation. CoSO_4 and CoF_2 obey the Curie-Weiss law. NiF_2 and NiBr_2 obey this law, but NiSO_4 shows such wide deviations from the linear function that a transformation of the cation at 280° is possible, raising the magnetic moment 2 units. However, the deviations for NiCl_2 do not permit such a conclusion. Perhaps here the equation $(\chi + K)(T + A) = C$ is applicable, where $\chi = \text{sp. susceptibility of the salt}$ and K the app. const. F. M. SYLVES.

Molar refraction of methanol. I. Influence of concentration of solution in a non-polar substance. M. VELASCO. *Anales soc. esp. fis. quim.* 29, 15-20(1931), cf preceding abstr.—The molar refractions of many solutions of MeOH in $\text{C}_{11}\text{H}_{22}$ were determined, showing that there is no effect of concentration on molar refraction. The bond between the dipoles of MeOH due to association does not modify the forces that act on the peripheral electrons of atoms constituting the molecule. E. M. SYLVES.

Dependence of the magnetic properties of the cobalt-chromium mixed crystals upon the temperature. I. RANZ WEVER AND HEINRICH LANGE. *Mitt. Kaiser-Wilhelm Inst. Eisenforsch. Düsseldorf* 12, 353-63(1930).—A series of Co-Cr alloys, in which the Cr content varied from 0 to 20%, was prepared from exceptionally pure materials. The magnetic properties of these alloys were studied over a wide range of temperatures with special attention to the region of the Co-rich mixed crystals. A description is given of the *astatic magnetometer* which was developed for this investigation and which proved to be very well adapted to tracing magnetization curves under unfavorable circumstances. The results are presented in a series of curves from which a number of conclusions are drawn. The polymorphic transformation of the Co-Cr mixed crystals shows an exceedingly large temperature hysteresis. In addition, to the retarding action of the Cr this seems to be due to the small energy difference between the types of crystal lattice when both are in equilibrium at the temperature of transformation. For both the modifications of the Co-Cr mixed crystals there is always a sharp temperature change for the loss of magnetic properties. These lines are very nearly parallel to each other. This confirms Heisenberg's deductions according to which the separation of the atoms and the coordination number determine the magnetic behavior. The absence of sharp breaks in the equilibrium curves for the polymorphic transformation again confirms the idea that the magnetic transformation is not to be considered as a phase transformation in the sense of the phase rule.

W. W. STIFLER

Change of the magnetic moment of nitric oxide with temperature. E. C. WIEBESMA, W. J. DE HAAS AND W. H. CAPEL. *Proc. Acad. Sci. Amsterdam* 33, 1110-24(1930).—The susceptibility of NO was measured in terms of the forces exerted on a small, evacuated glass bulb, placed in a non-uniform magnetic field, when surrounded by NO and by N_2 , resp. The diamagnetism of N_2 was assumed to be independent of the temperature. The results range from $\chi \times 10^4 = 49.07$ at 292.1°K to 87.32 at 112.8°K , and the agreement with the predictions of van Vleck's theory is excellent. W. W. STIFLER.

The magnetization-temperature curves of iron, cobalt and nickel. R. TYLER. *Phil. Mag.* [7], 11, 590-602(1931), cf C. A. 24, 4218.—The curves for σ/σ_0 against T/θ are almost identical for Fe, Co and Ni. They are definitely incompatible with the classical curve and with quantum curves for j greater than, or equal to, 1. There is close agreement when $j = 1/2$. The electron spin is to be regarded as the fundamental magnetic element operative for Ni, Fe and Co. The value of $j = 1/2$ means that electron spins are independent and that there is a $\tan \delta$ distribution. The electrons are not free but associated with ions as in Heisenberg's theory. ARTHUR FREISCHER.

Anomalous diamagnetism and crystal structure. V. J. VAIOYANATHAN. *Indian J. Physics* 5, 559-72(1930).—The anomalous diamagnetism of graphite, Bi, and Sb was found to be diminished on mechanical colloidalization. The decrease for graphite was greater than for the other two. Single natural crystals of graphite were found to be magnetically anisotropic, in the ratio of 1.67 for the crystal axis, resp., perpendicular and parallel to the field. In compounds of Bi and Sb the ionic values were found to be much smaller than for the crystalline state, e.g., for Sb^{III} , $\chi = -26 \times 10^{-6}$, for Sb^{IV} , $\chi = -66 \times 10^{-6}$, for Bi^{III} , $\chi = -40 \times 10^{-6}$, for Bi^{IV} , $\chi = -60 \times 10^{-6}$, and for

C in compds $\chi = -6 \times 10^{-4}$. The effects produced by colloidalization are ascribed to the existence of large electron orbits comprising groups of molts. Magnetically, amorphous C is considered to be microcryst. graphite and it is said that this view is supported by the x ray results of Delbye and Scherrer. The magnetic anisotropy of graphite, Bi and Sb is ascribed to the unequal valency groups for these elements.

W. W. STIFLER

Diamagnetism, field strength and crystal structure. W. J. DE HAAS. *Nature* 127, 335-6 (1931), cf *C A* 25, 855.—Some years ago de H. called attention to the probability of a correlation between crystal structure, change in elec. resistance in a magnetic field, diamagnetic susceptibility and possibly the Hall effect. Graphs are now given showing the susceptibility of Bi at 14.2°K. as a function of the field strength, both for fields parallel and perpendicular to the binary axis. Another set of curves shows the magnetization and change in resistance under the same conditions. In fields where the resistance increases, the magnetization remains nearly const. and vice versa. This is attributed to the behavior of the outer electrons and the explanation is offered that the diamagnetism of the outer electrons is superposed upon a paramagnetism due to the remainder of the atoms.

W. W. STIFLER

The diamagnetism of polyhalides. FRANCIS W. GRAY AND JOINT DAKERS. *Phil Mag* [7], 11, 81-95 (1931), cf *C A* 24, 1613.—The mol diamagnetic susceptibilities of alkyl and tri alkyl phenyl substituted N_2I_4 and Rb halides and polyhalides were detd. The structures are discussed by the method of Gray and Farquharson.

A. F.

Magneto-optical anisotropy in a plane normal to the optical axis of a hexagonal crystal. Paramagnetic rotatory powers and magnetic moments in directions close to the binary axis at very low temperatures. JEAN BECQUEREL AND W. J. DE HAAS. *Proc Acad Sci Amsterdam* 33, 937-44 (1930).—The magnetic rotations along the two varieties of binary axes differ and represent the first example of optical anisotropic phenomena in a plane normal to the optical axis of a uniaxial crystal. By the definition of a uniaxial crystal this cannot be true of a purely optical phenomenon, and it is actually due to the paramagnetic nature of the phenomenon and to the magnetic anisotropy. At the temps. close to 1.7°K. the ratio of the Verdet consts., 0.579 at 77.5°K., falls to 0.21.

A. P. SACHS

Paramagnetic rotatory power of a crystal of tysonite in a direction normal to the optical axis at liquid-hydrogen temperatures. JEAN BECQUEREL AND W. J. DE HAAS. *Proc Acad Sci Amsterdam* 33, 920-26 (1930), cf *C A* 25, 873.—The ratio of the magnetic rotatory powers along a direction making a small angle with the optical axis and along the optical axis, found equal to 0.579 at 77.5°K. (see preceding abstr.) decreases as the temp. falls.

A. P. SACHS

Magnetic rotatory power of a uniaxial crystal in directions oblique to the axis; determination of the rotation of tysonite in a direction close to a binary axis at the temperature of liquid nitrogen. JEAN BECQUEREL. *Proc Acad Sci Amsterdam* 33, 913-25 (1930).—Expts at ordinary temps. have already shown that tysonite has a smaller magnetic rotatory power oblique to the axis than along the optical axis (cf *C A* 23, 1792). At liquid N temp. the law of variation of magnetic rotation with the inclination of the wave and the direction of the magnetic field with respect to the crystallographic axes is confirmed. The observed magnetic rotation of tysonite in directions oblique to the axis is a paramagnetic rotation in the direction of a magnetization of the crystal resulting from the action of the magnetic field. The ratio of rotatory power normal to the axis to the rotatory power along the axis is 0.579 at 77.5°K., instead of 0.83 at 293°K.

A. P. SACHS

Determination of the susceptibility of cupric sulfate pentahydrate at low temperatures; the magneton numbers of the iron group. W. J. DE HAAS AND C. J. GORTER. *Proc Acad Sci Amsterdam* 33, 1101-5 (1930).—The susceptibility, χ , of $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ was measured at a no. of temps. from 290°K. down to 14.3°K. The results range from 6.95×10^{-4} at room temp. to 121.6×10^{-4} at the temp. of liquid H_2 . When corrected for the diamagnetism of the anion and of the H_2O of crystal, the values are $\chi' = 8.35 \times 10^{-4}$ and 1.22×10^{-4} , resp. The results follow the Curie-Weiss law with $\theta = -70^\circ$ and $\chi' T = 18.3$. The corresponding magneton no. is 9.51 Weiss magnetons. In the light of these results, the theory of Bose and Stoner is extended to cover deviations in the second part of the Fe group and an explanation is offered for the variable magneton nos. found there.

W. W. STIFLER

Contribution to the theory of dielectrics. G. GUÉZENY. *Phil Mag* [7], 11, 405-10 (1931).—Math. A formula is derived on the assumption based on expl. facts that the current in a dielec. is due both to the contribution of dipoles and to the presence of ions.

ARTHUR FLEISCHER

The dielectric constant of water as determined by a resonance method. E. P. LINTON AND O. MAASS. *J. Am. Chem. Soc.* 53, 957-64(1931).—The resonance method of Cuthbertson and Maass (*C. A.* 24, 1775) was reinvestigated and a method for calibrating condensers, which vary linearly according to their setting, is described. The effect of power of oscillating circuit and precautions required in work with media of high dielectric const. are given. The effect of cond. of medium on resonance point is negligible provided the cond. is below 4×10^{-4} (in the case of water). The dielec. const. of water is 79.2 at 25° based on 1.00 for air.

Dielectric characteristics of Seignette's salts. P. P. KOREKO AND I. V. KURCHATOV. *J. Russ. Phys.-Chem. Soc., Phys. Pt.* 62, 251-65, *Z. Physik* 66, 192-205(1930).—Experimentation indicates that high values of dielec. const. cannot be explained by high polarization. The increase in the induction is believed due to (1) transmutation of electrons into atoms, (2) vibration of ions within the lattice and (3) orientation of mols. possessing a dipolar moment.

New results on dielectric constants of electrolyte solutions (from measurements by A. Slama and O. Milicka). R. FORTI. *Physik Z.* 32, 184-7(1931), cf. *C. A.* 18, 1606; 19, 1080, and Pechold, *C. A.* 22, 907.—It is shown that dielec. const., D , determined by the ellipsoidal method of I. is independent of the frequency, f , between 50 and 2200 cycles/sec. The method is improved to give a precision of 0.6% up to a cond. of 10^{-2} ohm $^{-1}$ cm $^{-1}$. The curves obtained by plotting D against c (molar concn.) are all of the same type: with increasing c , D decreases to a min. (at $c = 0.005-0.01$), then increases abruptly and finally approaches a limiting value which is frequently above the D of H_2O . With electrolytes of the 1-2 or higher valence types, the c, D curve may show two min. or an inflection and a min. The min. D shifts toward higher c with increasing f . The lowering of D observed here is much greater than at higher f (e.g., 10^4 cycles/sec). In the c range near the min. D (where the electrolytes are largely dissociated), it is assumed that most of the H_2O mols. are held oriented in the H_2O sheaths about the ions. Then as c increases, the ions tend to unite forming complexes and liberating H_2O mols. which are now free to rotate in accordance with an external field and D therefore increases. The various min. or inflection points probably indicate the formation of one or another complex. Upon the assumption that, at the c of min. D , all H_2O is bound by the ions, m (the no. mols. H_2O per mol. of electrolyte) and n (the no. mols. H_2O per unit ionic charge) were calcd. for $HCl, LiCl, NaCl, KCl, FeCl_3, CuCl_2, BaCl_2, PbCl_2, AlCl_3, FeCl_3, NaBr, NaI, HNO_3, NaNO_3, Ba(NO_3)_2, NaOH, KOH, Ca(OH)_2, H_2SO_4, H_3BO_3$ and H_3PO_4 . The values of m varied from 125 to 385 and those of n from 2×10^4 to 2×10^5 . In general, n increased with the wt. of the anion (resp., cation) in a series of electrolytes having a common cation (resp., anion). This is in accord with the idea that the larger the ionic wt. the more points of attachment the ion would have for H_2O mols. The H_2O sheath is pictured as being formed of more or less oriented H_2O dipoles, the attractive force being "passed on" from link to link of the chain as in the case of iron filings picked up by a magnet. At low f the time required for the formation of such a "chain-model," H_2O sheath is assumed to be small with respect to the duration of an oscillation but very large with respect to the relaxation time of the H_2O dipole. It is assumed further that at $f = 10^4$ this chain formation is too slow to take place and thus at high f more H_2O dipoles are free for orientation in an external field; the D lowering is therefore smaller.

Molar refraction of methanol. M. VELASCO. *Anales soc. españ. fis. quim.* 28, 1228-38(1930).—The sp. inductive power of solns. of $MeOH$ in C_6H_6 was detd., and the results were used to calc. the molar polarization of the solns. Variation of molar polarization of $MeOH$ as a function of the molar fraction contained in the soln. was examined by the Debye theory to deduce the molar polarization at infinite diln. By the aid of this limit value the dielec. moment of the $MeOH$ mol. is calcd. as $\mu = 1.6 \times 10^{-18}$.

Dipole moments of some organic sulfur compounds and related substances. III. Meaning of dipole measurements in the stereochemistry of carbon compounds. ERNST BERGMANN, LEO ENGEL AND STEFAN SANDOR. *Z. physik. Chem., Aht. B.* 10, 397-413(1930); cf. *C. A.* 24, 4274; 25, 212.—The following moments are given: diphenyl sulfide 1.47, *p*-chlorodiphenyl sulfide 1.38, thioanisole 1.27, diphenyl selenide 1.38, diphenyl telluride 1.14, phenyl thiocyanate 3.59, *p*-chlorophenyl thiocyanate 2.93, benzyl cyanide 3.91, *p*-bromobenzyl cyanide 2.64, *o*-chlorobenzyl cyanide 4.76, phenyl mustard oil 2.76, diphenyl sulfoxide 3.83, nitrobenzene 4.03, fluoronitrobenzene 2.63 and benzyl chloride 1.85. From these values are established the characteristic valence angles for the following groups: SC_6H_5 , 71°, S , 142°, SCN , 127° and CN , 154°. *p*-Substitution has little effect upon such angles, but they are affected by the susceptibility to polariza-

tion of the mol. as a whole. The mustard oil group is built at an angle. The earlier mentioned "semi polar double bonds" cannot be confirmed. A model for the last 2 compds. is described from which the moments can be calcd. with satisfactory accuracy.

GREGG M. EVANS

Comments on the paper by Shilling and Larton on "The effect of temperature on the viscosity of air." A. O. RANKINE. *Phil Mag* [7], 11, 225-7 (1931); cf. *C. A.* 25, 10—Polemical.

ARTHUR FLEISCHER

Flame temperatures of hydrocarbon gases. G. W. JONES, BERNARD LEWIS, J. B. FRIATY AND G. ST. J. PERROTT. *J. Am. Chem. Soc.* 53, 809-83 (1931)—Measurements of flame temps. for air mixts. are reported for the following gases. Pittsburgh Natural Gas (CH_4), C_2H_6 , C_3H_8 , butane, isobutane C_4H_{10} , C_5H_{12} , and C_6H_{14} . A spectral line reversal method with Na and Li salt sprays was employed. The temps. are a max. when the mixt. has a slight excess of combustible gas over the amt. necessary to consume all the O_2 present. These max. temps. range from 1880° for CH_4 to 1975° for C_6H_{14} and agree with calcd. values. Max. flame temp. is attained with a mixt. contg. less combustible gas than the mixt. that gives max. speed of uniform flame movement.

J. B. AUSTIN

The velocity of sound in carbon dioxide. H. O. KNEZER. *Physik Z.* 32, 179 (1931)—The velocity of sound in CO_2 increases with the frequency in the range 0.5– 3.0×10^3 cycles/sec. Above the frequency 3×10^3 the velocity is again const. at 268.2 m./sec. (3.7% higher than the normal value). This leads to a value of 1.40 for c_p/c_v .

OSCAR T. QUINBY

A relation between orthobaric volumes and temperature. JERRO HORTON HARRIS. *Sci. Papers Inst. Phys. Chem. Research* 15, 89-104 (1931); cf. *C. A.* 21, 860, 3871.—The applicability of H.'s equation of orthobaric vol. to CH_4 , C_2H_6 , C_3H_8 , C_4H_{10} , MeCl , EtCl , Me_2O , CO , Me_2S , Et_2S , PH_3 , Cl_2 , HCl , H_2 , Ne and Xe is demonstrated. The various parameters are tabulated for the 58 substances studied in this and the previous papers. The rules and approximations derived from the H. equation were used to calc. the orthobaric d. of Kr as follows: 210° , 0.92 and 0.82; 200° , 1.34 and 0.40; 190° , 1.58 and 0.25; 180° , 1.75 and 0.17; 170° , 1.89 and 0.11; 160° , 2.00 and 0.07; 150° , 2.10 and 0.04.

OSCAR T. QUINBY

Properties of saturated vapors, including steam. A. FRESS. *Phil Mag* 11, 111-8 (1931)—Math. The first law of thermodynamics is solved to give 2 functional equations. The first was used by Callendar for the construction of steam tables, while the second leads to a theoretical justification of the third law. For satd. vapors, the equation of state takes the form $p = (F - 1) a_m / (V - F)$, where F is a const. characterizing a fluid and is the ratio of the gross latent heat to the net latent heat of evapn. For steam from 0° to 220° , $F = 1.11$.

ARTHUR FLEISCHER

The pressure-volume-temperature values for ammonia to one thousand atmospheres from 30° to 200° . FREDERICK G. KEYS. *J. Am. Chem. Soc.* 53, 965-7 (1931), cf. *C. A.* 12, 550.—The P - V values for NH_3 (with vols. of 1.50–5 cc./g.) were detd. from 30° to 200° with a max. error of 0.25%.

W. C. FERNELITS

The equation of state of propellant gases. A. D. CROW AND W. E. GRIMSHAW. *Trans. Roy. Soc. (London)* A230, 39-73 (1931)—The app. and exptl. procedure are described for measuring the pressures developed by exploding nitrocellulose, cordite and ballistite in a vessel of given capacity. Observations were made over a range of loading densities and for 2 vessel capacities, the observed pressures being corrected for the cooling of the hot gases caused by the wall surface of the vessel and the work absorbed in the wall stress. The equation of state of the hot gaseous mixt. resulting from the explosions is the expression for the pressure exerted by the uncooled gases and is found to be $p_0 = \lambda / [(1/\Delta) - \eta]$, where λ is the ballistic force of the propellant, Δ is the density of loading, and η is defined as the co-volume, its numerical value being about 4 times that of the mols. contained in 1 g. of the gas complex. The expression for the ballistic force is $\lambda = 84.80 \{ (2 \text{ g. mols.}) / g \} T_0$, where T_0 is the temp. of the explosion and depends on the constituents of the explosion, their heats of formation, the mean mol. heats at const. vol. of the gases over the temp. range involved (2700°K. to 4200°K.), and the control equation of the water-gas reaction. The co-vol. η is evaluated from the constituents of the gas complex by using the hard kernel values of the mol. radii increased by 25%.

C. C. KREBS

The constants of the Beattie-Bridgeman equation of state with Bartlett's P - V - T data on hydrogen. W. EDWARDS DEXING AND LOLA E. SHUPPE. *J. Am. Chem. Soc.* 53, 843-9 (1931), cf. *C. A.* 24, 2925.—Applying Bartlett's P - V - T data (*C. A.* 22, 2302) to the Beattie and Bridgeman equation of state, $P1^2 = RT(V + B_0(1 - b/V))(1 - c/V^2) - A_0(1 - a/V)$, gives the consts. for H_2 . Below the crit. d., $A_0 = 124.040$, $B_0 = 20.22$, $a =$

56.18, $b = -7.22$, $c = 20 \times 10^4$, above the crit. d. $A_0 = 124.040$, $B_0 = 17.50$, $a = 56.18$, $b = -19.68$, $c = 20 \times 10^4$. A break in the graph of B vs d occurs near the crit. d. The Beattie-Bridgeman equation of state and Bartlett's $P \cdot V \cdot T$ data on a 3:1 hydrogen-nitrogen mixture. *Ibid* 860-9.—For a 3:1 mixt. of H_2 and N_2 , the consts. are: Below the crit. d., $A_0 = 3000 \times 10^4$, $B_0 = 23.03$, $a = 21.36$, $b = -15.16$, $c = 16 \times 10^4$, above the crit. d., $A_0 = 3489 \times 10^4$, $B_0 = 21.42$, $a = 28.56$, $b = -30.58$, $c = 16 \times 10^4$. A break in the graph of B vs d , again occurs near the crit. d.

DON BROUZE

State equation of easily liquefiable hydrocarbons. IV. Weight of the liter and compressibility coefficient of normal butane. MARCEL BECKERS. *Bull. soc. chim. Belg.* 39, 470-93 (1930), cf. *C. A.* 24, 761.—This is a very careful exptl. detn. All the precautions needed to obtain the high degree of accuracy claimed are extensively described. It was found that for C_4H_{10} , $(\rho, \kappa) = 0.997013$. The wt. of the normal l. is 2.70324.

A. L. HENVE

Dimensional analysis applied to the thermal conductivity of liquids. J. F. DOWNS SMITH. *Ind. Eng. Chem.* 23, 416-9 (1931).—Thermal cond. of liquids is assumed to be a function of mol. wt., d , ρ , heat, viscosity, gas const., thermal expansion and compressibility. By dimensional analysis, a formula is obtained connecting all these variables. It may be simplified and 2 functions plotted against each other, the graph showing a max. error of 4.5% in the thermal cond.

G. M. MURPHY

Prediction of critical temperatures and heats of vaporization. K. M. WATSON. *Ind. Eng. Chem.* 23, 360-4 (1931).—The crit. temp. of a non polar liquid may be calc'd from its b. p., mol. wt. and liquid d , with an error of less than 2%. The latent heat of vaporization at any temp. or pressure may be calc'd. by 2 methods, when the crit. temp. and normal b. p. are known. The error is generally less than 5%. For either polar or non-polar liquids, the latent heat of vaporization may be calc'd at any temp. if the crit. temp. and 1 exptl. value of the latent heat are known.

G. M. MURPHY

Allotropy in liquids. L. A. SMITS. *Z. physik. Chem., Abt. A*, 153, 287-98 (1931), cf. *Verslag Akad. Amsterdam* 1912, 418.—The allotropic theory is extended for the cases of monotropic and enantiotropic liquid transitions.

ARTHUR FLEISCHER

The influence of intensive drying on inner equilibria. L. A. SMITS, E. L. SWART, P. BRUN AND W. M. MAZER. *Z. physik. Chem., Abt. A*, 153, 255-61; *Proc. Acad. Sci. Amsterdam* 34, No. 1, 160-4 (1931); cf. *C. A.* 21, 1049, 22, 1523, 23, 1827.—Eight mos.' intensive drying over P_2O_5 in a modified app. (cf. Terwen, *C. A.* 11, 553) insuring a gas-free liquid showed no effect in accord with those recorded by Baker. Baker's effects on intensively dried liquids were due to superheating which occurs readily in dust free liquids.

ARTHUR FLEISCHER

Superheating and intensive drying of liquids. ERNST COHEN AND W. A. T. COHEN-DE MEESTER. *Proc. Acad. Sci. Amsterdam* 33, 1003-14 (1930); *Z. physik. Chem., Abt. A*, 153, 241-53 (1931).—The authors confirmed the work of Lenher (cf. *C. A.* 23, 4388) and concluded that the abnormal rise in the temp. of the boiling liquid observed by Baker (cf. *C. A.* 16, 2441) was not due to the drying.

GREGG M. EVANS

Apparatus and methods for precise fractional distillation analysis. New method of gas analysis. WALTER J. PODBIELNIAK. *Ind. Eng. Chem., Anal. Ed.* 3, 177-88 (1931); cf. *C. A.* 24, 5992.—The construction and operation of the latest development in fractional distn. of complex hydrocarbon mixts. are described. The operation is based on distn. at low temps. and at pressures of 1 atm. abs. or below. Distn. curves are given for natural gas, manuf'd. gas, refinery tail gases, natural gasoline and other hydrocarbon mixts. having compds. boiling below C_4H_{10} . The accuracy of analyses is discussed.

M. C. ROGERS

Measurement of viscosity and its applications in chemistry. A. BOUTARIC. *Industrie chimique* 17, 762-5, 841-3 (1930), 18, 6-7 (1931).—A review.

P. T.

Viscosity of liquids. Methods of measurement and applications. A. BOUTARIC. *Rev. chim. ind.* 39, 194-9 (1930).—A review.

P. THOMASSET

Hydrodynamics of systems of varying viscosity. IV. MARCUS REINER. *Kolloid-Z.* 54, 175-81 (1931), cf. *C. A.* 24, 2652.—The fundamental law $\eta = \eta_0 - (\eta_0 - \infty)e^{-\epsilon/\epsilon_0}$ is derived for non-Newtonian liquids, ϵ , liquids with structural viscosity. In the equation ϵ_0 is the fluidity at rest, ϵ , at the largest fluidity, C the stability coeff. of the liquid structure, ϕ the fluidity at the tangential tension τ . An equation for the capillary viscometer is also derived.

ARTHUR FLEISCHER

Necessity of using the absolute kinematic viscosity coefficient in the practical determination of viscosity. PAUL WOOD. *Bull. inst. phys.* 1931, 23-4, 41-7.—After showing the function and importance of viscosity and discussing the effects of temp. thereon, the chief viscometers now in use and the various methods available for actually measuring

the true viscosity coeff are described. A plea is made for the adoption of such methods and of the poise and stokes as practical units of measurement. A. P. C.

Experimental contribution to study of uniform laminary flow. G. B. UCOLINI. *Ann. lavoro pubblico* 68, 490-510 (1930).—The generally assumed parabolic distribution of velocity along the tube diam is not substantiated by expts with water in cylindrical glass tubes 4.5 mm in diam. Velocity distribution seems to be more like that of a turbulent stream. E. I. S.

The mechanism of plastic flow. G. F. CUNNINGHAM. *J. Phys. Chem.* 35, 700-814 (1931).—The mobility of clay pastes was found to vary markedly with the pressure. L. ONSAGER.

The surface tension of crystals. D. HALASZEV. *Kolloidchem. Beihefte* 32, 205-11 (1931).—No equil can exist between a single crystal and its mother liquor for a homogeneous surface according to the Ostwald-Landlich equation. Such an equil is possible provided the surface tensions at edges and corners are greater than the surface tension at a face. The faces are believed to grow in a mosaic pattern thus giving an inhomogeneity to account for a greater surface tension. R. H. LAMBERT.

Strength of salt crystals partially immersed in water. A. SNEKAL. *Physik Z.* 32, 187-92 (1931), cf. *C. A.* 23, 3836.—The tensile strength of NaCl crystals, whose mid-ports were bathed in water, was much greater than that of a normal dry crystal, but the increase was less than a tenth that reported by Joffé (*C. A.* 22, 518). OSCAR T. QUMBY.

Progress and problems of chemical crystallography. H. STEINMETZ. *Fortschr. Min. Arist. Petr.* 12, 193-219 (1927). E. H.

Some examples of cyclic twin formations. R. BRAUNS. *Centr. Mineral. Geol., Paldont*, Abt. A, 1929, 357-66; *Chem. Zentr.* 1930, I, 957.—Cyclic twin formations, which give the impression that the crystals have a higher symmetry than they possess in reality, are described with K_2SO_4 and alexandrite as examples. With the former the twin planes are (110) and (110) and with the latter they are (130). Serial twins were not observed on either one. G. SCHWOCIT.

The possible orientation of cubic crystals deposited on a sheet of mica. L. ROYER. *Compt. rend.* 191, 1346-8 (1930).—Crystals of alk. fluometaphosphates are oriented on mica similarly to those of alk. halides, suggesting an analogy between the structures of the two series of compounds. Orientation takes place only if the structure is octahedral and of the face-centered cubic class like NaCl. Values for the unit length of the (111) face of KPF_6 , $RbPF_6$, NH_4PF_6 , and $CsPF_6$ are calcd and it appears that the difference between these values and that for mica (5.12 Å) in a case of mutual orientation of crystals of different species must not exceed 0.69 Å. H. W. WALKER.

The dehydration of heulandite studied by means of x-rays. J. WYART. *Compt. rend.* 191, 1343-6 (1930).—Up to about 210° the axis of symmetry of heulandite contracts 0.35 Å or 2% with loss of water. No modification in the at arrangement takes place and the change is reversible. Above 210° with a 10% loss in wt a large contraction occurs, the crystal form is altered, and the structure becomes heterogeneous like that of dehydrated mica and kaolinite. Cryst. platelets of diverse orientation result. Such a change is irreversible. H. W. WALKER.

Crystallographic researches on some racemic compounds and the respective optical antipodes of the asparagine group. GUIDO CASORRI. *Rend. accad. sci. Napoli* 36, 49-55 (1930).—The following properties were detd: *L*- and *D*-Chloroacetylparaspargine: m. 150°, monoclinic, sphenoidal, $a \ b \ c = 0.7876 \ 1.04281$, $\beta = 96^\circ 22'$; simple forms $a(100) \ b(010) \ b'(010) \ m(110) \ m'(110) \ \gamma(101) \ \gamma'(101) \ \gamma(121) \ x(121)$; elongated along the *y*-axis, the bisectrix emerges from (100), inclined to the normal of that plane, and at a large angle relative to the optic axes, plane of the optic axes is normal to (010), sp gr 1.538, mol vol 135.59. *DL*-Chloroacetylparaspargine: m. 105°, triclinic, pinacoidal; $a \ b \ c = 0.9068 \ 1.08617$, $\alpha = 63^\circ 23'$, $\beta = 113^\circ 20'$, $\gamma = 117^\circ 15'$; simple forms, $a(100) \ b(010) \ c(001) \ n(110) \ m'(110) \ \gamma(101) \ \gamma'(101)$, elongated along the *y*-axis and somewhat tabular relative to (100); in convergent light one axis emerges from (100) at the edge of the field, on (100) the direction of max extinction makes an angle of 5° with the *y*-axis in the obtuse angle formed by [110.100] and [001.100]; on (001) the extinction angle is 22° with the *y*-axis in the obtuse angle formed by [001.100] and [001.110], sp gr 1.560, mol vol 145.23. *L*- and *D*-Benzenesulfonylparaspargine: m. 163°, monoclinic, sphenoidal, $a \ b \ c = 0.7557 \ 1.07105$, $\beta = 104^\circ 42'$; simple forms, $a(100) \ b(010) \ b'(010) \ c(001) \ m(110) \ m'(110) \ g(011) \ g'(011)$, often tabular with respect to (100), in convergent light, crossed nicols, an optic axis emerges from (100) at the edge of the field, plane of the optic axes parallel to (010), sp gr 1.434, mol vol 189.80. *DL*-Benzenesulfonylparaspargine: m. 172°, minute crystals, mono-

cubic prismatic, elongated relative to the z -axis, and tabular relative to a pair of parallel faces, probably (100). In convergent light, with the crystal lying on (100), there appears one bisectrix slightly inclined, and both optic axes at the edge of the field, plane of the optic axes is normal to (010), σ gr 1.559, mol vol 171.50. *dl-Benzoylparaphenyl* monochloride, prismatic, a b c = 0.8805 1.05020, β = $92^\circ 12'$; simple forms, $a(100)$ $b(010)$ $c(111)$ $a(111)$, elongated along a and tabular relative to (100), on (010) a direction of max extinction makes an angle of 7° with the z -axis in the angle formed by [101 100] and [100 111], the acute bisectrix emerges slightly inclined from (100) at a small angle with the optic axis, the plane of the axes is normal to (010), angles of the optic axes are $21^\circ 12'$, $21^\circ 54'$, $28^\circ 50'$, $21^\circ 12'$, $30^\circ 10'$, σ gr. 1.433, mol vol 161.70. Values of the interfacial angles are given for the above crystals.

R. H. LAMBERT

Regularity in crystal structure in hydrides, borides, carbides and nitrides of transition elements. GUNNAR HÄGG *Z. physik. Chem.*, Aht B, 12, 33-56(1931); cf. *C. A.*, 24, 1591.—The structure of crystals of binary compounds containing a so-called "transition element" and either H, B, C, or N is critically examined with respect to the ratio of atomic radii. If this ratio $r_m/r_e > 1.7$ the structure is simple, while if the ratio is smaller the structure becomes more complex the smaller the radius of metal ion. Only these combinations show a typical metallic character in crystal structure which is termed an infiltration structure. The latter shows a combination M_2X , M_3X , MX and MX_2 , each of which is described. The face centered cubic or close packing hexagonal lattice, both of coordination no. 12 are most often observed. One case each of body-centered cubic and simple hexagonal, both of coordination no. 8 was found. A careful discussion follows of the packing arrangement of atoms and the relation to coordination nos.

R. H. LAMBERT

Röntgenographic investigations of potassium fluoborate. B. PRSEN *Gazz. chim. ital.*, 60, 936-9(1930).—Bacchi and Mark found that $SeSO_4$, $FeSO_4$, $KMnO_4$ and $KClO_4$ are of the same structural type, and contain 4 mols in the elementary cell (cf. *C. A.*, 20, 3597). KBF_4 should belong to this same series, for it is isomorphic with $KClO_4$ and $KMnO_4$ (cf. Zambonini, *C. A.*, 17, 2000). It was therefore of interest to study KBF_4 by x rays (Cirr radiation). The exptl. d_{111} was 2.551-559 (cf. d_{111} 2.198 of SnF_4 , *Jahresber. Chem.* 1872, 231). The x-ray examn. showed $a_1 = 7.81$, $b_1 = 5.68$, $c_1 = 7.38$ (rhombic lattice), with a b c = 1.382 1.1297. Montemartini reported the prepn. of a cubic form of KBF_4 (cf. *Rend. accad. Lincei*, 3, 549(1891); *Gazz. chim. ital.*, 24, 1, 478(1894)), but by following the procedure of M., a product was obtained that gave röntgenographs identical with those of the rhombic form, whether the gelatinous compound was dried in air without heating or dried at 100° . After drying at 100° , its d was 2.555. The so-called cubic form and the cryst. rhombic form are therefore identical.

C. C. DAVIS

An x-ray study of mannitol. G. W. McCREA, *Nature*, 127, 162-3(1931).—Results for d -mannitol are space group C_2 , $a = 8.64$ A. U., $b = 10.58$ A. U.; $c = 5.50$ A. U., calcd 1.522, 4 mol. per cube.

MALCOLM DOLLE

Conductivity of zirconia. HANS SCHWITZER *Z. anorg. Chem.*, 44, 151-2(1931).—The cond. was detd. between 200° and 800° (150-1100° abs.). One sample of ZrO_2 contained 2% of MgO and the other $1\frac{1}{2}\%$, a trace of TiO_2 and 0.3% H_2O . Graphs are given showing the log cond. against the reciprocal of the abs. temp. The curves for both the specimens show breaks, approx. at $450-550^\circ$, when the values obtained with increasing temp. are plotted. Values obtained with decreasing temp. gave a straight line considerably below the other values. It is suggested that this is due to a loss of H_2O which occurs in the pure material at approx. $510-50^\circ$. The purer sample (contg. no Mg) showed a much higher cond. and it is reasoned from this that the cond. of fused, pure ZrO_2 is probably appreciably smaller than that of quartz. The relationship between cond. and field strength for the Mg -free ZrO_2 , which had been heat-treated at 800° , was found to be linear up to 1500 v./cm.

EDWARD B. SANGAR

The properties of chlorides of sulfur. V. Metastable states. THOMAS M. LOWRY AND GILBERT JESSOP, *J. Chem. Soc.* 1931, 323-9; cf. *C. A.*, 24, 4111.—The period of induction in the combination of SCl_2 with Cl_2 can be eliminated or reduced by catalysts such as I and $SbCl_3$ or by autocatalysts. The combination of SCl_2 and Cl_2 to form SCl_4 is instantaneous even at -80° , no lag can be detected in restoration of equl. when SCl_2 is melted below -20° . The freezing and melting of SCl_4 is characterized by a marked lag. A displacement with temp. of the equl. concns. in samples of SCl_4 was established by observations of d , the disson. increases from about 15% at 0° to about 25% at 100° .

R. H. LAMBERT

Langmuir's adsorption isotherm. R. S. BRADLEY *Phil. Mag.*, 11, 600-6(1931).—

Polanyi's theory is criticized on the basis that the heat of adsorption and the adsorption potential cannot be identified. The capillary condensation method is inconclusive. By the application of the theory of imperfect gases in a two-dimensional phase, the equation, $\log c_1/c_2 = \text{const}/T + \text{const}$, is derived. The const. c_1 and c_2 are those of Langmuir's equation. The equation was found to apply except at low temp.

ARTHUR FLEISCHER

Adsorption intensity and its importance in technical processes. ISIDOR TRAUBER *Z. anorg. Chem.* 44, 73-8(1931).—Several means of measuring adsorption intensity have been, or are being developed. Where no reaction occurs the intensity of adsorption is proportional to the amt. of heat generated, the phys. conditions of the adsorbant and the substance undergoing adsorption being always a const. Methods based upon displacement principles and vapor pressure were also found satisfactory for estg. adsorption. The adoption of such a standard method would be of great value in the *flokation and lubricating-oil industries* and be an aid to botanical studies.

W. H. TIFFANY

Some relationships between phase-boundary potential, adsorption, surface tension and particle size. LAURENCE K. V. KOCH *Phil. Mag.* 11, 585-62(1931).—Math. The equation for the difference in potential between a spherical particle and a plane surface is derived.

ARTHUR FLEISCHER

Discontinuous nature of the process of sorption of gases and vapors by porous solids. A. J. ALLMAND AND L. J. BURRAGE *Proc. Roy. Soc. (London)* A130, 810-22(1931).—Work pointing to the discontinuous nature of the sorption of vapors by charcoal is reviewed. Eight isothermals of the sorption of CS_2 , CCl_4 , C_6H_6 , H_2O , $\text{C}_2\text{H}_5\text{OH}$ and CO_2 on different charcoals showing the great no. of observations made are given. Isothermals of C_6H_6 , CCl_4 and H_2O on silica gel showing the discontinuous sorption are presented. The theory of the sorption of vapors over a wide range of concn on charcoal is discussed.

S. LEVINE

Heats of wetting and of adsorption on zinc oxide. WARREN W. EWTING *Ind. Eng. Chem.* 23, 427-9(1931).—Heats of adsorption and of wetting of CO_2 , NH_3 , SO_2 , Nujol, C_6H_6 , $\text{C}_2\text{H}_5\text{OH}$, $\text{C}_6\text{H}_5\text{N}$, linseed oil, 15% Zn oleate in C_6H_6 , 5% Zn oleate in $\text{C}_6\text{H}_5\text{N}$ and H_2O on ZnO have been measured at 25° with a simple calorimeter.

S. L.

Specific adsorption properties of activated carbon. L. M. M. DUBININ *J. Russ. Phys.-Chem. Soc.* 62, 1627-33(1930); cf. *C. A.* 25, 13.—A preliminary theoretical discussion of possible properties of activated carbon to be followed by exp't. investigations. The quantity of adsorbed gas is expressed by $a = Cr \times S$ (Cr is the quantity of gas adsorbed by one unit of adsorbent surface and the surface S depends upon the diam. of the pores D , $S = f(D)$). In case of adsorbed liquids the activity is characterized by the sum of volumes of the pores and their radius $r = e(S)$. In case of solns. the equation for the adsorption of gas $a = Cr \times S$ may be used. Cr depends upon the nature of the solvent, nature of the surface of the carbon (amorphous, cryst., interacted with O of the air, contg. impurities, formation of an elec. charge greatly affects adsorption of strong electrolytes). The value of S (adsorbing surface) depends upon the nature of the adsorbed mols. of the solute.

J. G. TOLPIN

The adsorption of ethylene on wood charcoal, silica gel and alumina. A. MAGNUS AND H. WYDECK *Z. physik. Chem., Abt. A*, 153, 113-26(1931); cf. *C. A.* 23, 2340, 24, 4683.—Adsorption curves were det'd. from low pressures to about 550 mm. at 0°, 20°, 40° and 60°. Comparison at 0° and 50 mm. pressure shows that 1 g. of the beechwood charcoal, 3.6 g. SiO_2 gel, and 5.33 g. of Al_2O_3 adsorb equal quantities of C_2H_4 . The adsorption curves for charcoal show a sharp rise against the flat curves shown by SiO_2 gel and Al_2O_3 . The value of β/θ for charcoal decreases with temp. although the theory requires it to be const. For SiO_2 gel and Al_2O_3 fictitious values of β/θ , a term in the Magnus equation of state for adsorbed gases, are obtained.

A. F.

Adsorption within the crystal of a salt. I. N. STRANSKI. *Kolloidchem. Beihefte* 32, 197-204(1931).—S. does not believe that the Gibbs-Ostwald-Freundlich equation relating soly. and surface tension of particles is sufficient to explain the effect of grain size. He objects to Bairew's freely interchanging the terms surface tension and soly. Various extracts from B's paper are given and criticized. The assumption that water of hydration for crystals like those of BaSO_4 is "free" seems untenable. The theoretical results given by B. therefore can lead to erroneous conclusions.

R. H. L.

The kinetics of surface processes on crystal lattices. I. The adsorption system barium sulfate-electrolyte solutions. LUDWIG ICKA *Z. physik. Chem., Abt. A*, 153, 262-86(1931), cf. *Z. physik. Chem., Abt. A*, 153, 127(1931).—The adsorption of radioactive Th B, isotope of Pb and of Act, isotope of La, was det'd. on pptd. BaSO_4 at 18° in the presence of various ions. The velocity of adsorption of Act is decreased by the addn. of Ba ions and increased by the addn. of SO_4 ions. Foreign electrolytes such as

HCl diminish the speed of adsorption. Qualitatively the curves for Th B are similar. Quant inspection reveals the differences that Th B is adsorbed to a much smaller extent and that the initial slope of the Act curves is much greater than that of the ThB curves. The differences illustrate the previous results. The Act curves are characteristic of those of easily sol electrolytes in which valence is the important factor. The Th B curves are characteristic of those of difficultly sol electrolytes in which the velocity is controlled by the recrystn. of the adsorbent in accommodating the exchange reaction.

ARTHUR FLEISCHER

Adsorption. JOHANNES TERWELLEV *Z physik Chem*, Abt. A, 153, 52-67 (1931).—The adsorption of methyl violet and methylene blue on Ag and Ni wires was detd. by measuring the change in concn. of the solns spectrophotometrically. To avoid aging effects in the dye solns, pure dyes were used, the solns were kept in the dark except for the photometer measurements, and O_2 and CO_2 were excluded from the solns. The curves for amount adsorbed against dye concn. rise rapidly to a satn. value, which remains const. up to a high concn., after which the adsorption again increases rapidly. Methyl violet on Ag shows a bimol. layer if a cubic mol. is assumed, and a monomol. layer if the mol. has a cross section of 3.263×10^{-15} cm.² and a length of 17.49×10^{-8} cm. Methylene blue on Ag is adsorbed as a monomol. layer if the cubic form of the mol. is assumed. Heating the Ag wires to 400° for 2 hrs reduced the adsorption considerably. Methylene blue is adsorbed more readily by Ni than by Ag, while methyl violet is adsorbed to a smaller extent. Electrolytic Ni (marked II) adsorbed more than the regular Ni wire. The satn. values in g $\times 10^4$ for 1000 cm.² of wire surface are given in the table.

	Ag	Heated Ag	Ni I	Ni II
Methyl violet	2.092	0.694	1.263	3.633
Methylene blue	1.078		1.877	1.927

The satn. value depends on the nature of the adsorbent and on the dyestuff so that detn. of surface by detn. of the adsorption is an empirical method. Langmuir's theory of monomol. adsorption is not confirmed in the above results. ARTHUR FLEISCHER

Adsorption. HERMANN SCHLICHTER *Z physik Chem*, Abt. A, 153, 68-82 (1931).—The adsorption of CS_2 from 0.05 to 0.9 p/p. was detd. on glass powder at 43.5° and on Ag powder at 25° and 43.5° . The adsorption of pentane on the two powders was detd. at 34° . The curves are all similar and may be expressed by the equation $N = a b^P$, where N is the no. of layers of mols. in the adsorbed film, P is the pressure and a and b are const. In all cases the adsorbed layer is polymolecular. The results can not be explained by capillary adsorption or by a soln. effect like that of water on a glass surface.

ARTHUR FLEISCHER

The activity coefficients and the adsorption of organic solutes. I. Normal butyl alcohol in aqueous solution by the freezing point method. WM. D. HARRIS AND ROY W. WAMPLER. *J. Am Chem Soc* 53, 850-9 (1931).—The activity values for the aq. solns. of BuOH were detd. by the use of a specially designed app. for the detn. of f.p. lowering sensitive to 0.00001°. The activity coeffs. of the carefully purified BuOH in 1, 0.5 and 0.1 M solns. are 0.8227, 0.8681 and 0.9433, resp. Calcns. based upon the usual "adsorption equation" show that the film of BuOH on aq. solns. is monomolecular; the area per mol. is 237 sq. Å. U. P. H. EMMETT

The effect of the hydrogen-ion concentration in the adsorption of ions of barium, aluminum and thorium by clays. N. A. HELD AND M. N. SOKOLOVA. *J. Russ. Phys.-Chem Soc* 62, 1571-5 (1930).—Glukhov kaolin adsorbs Ba-ion from solns. of $Ba(OH)_2$ to a much larger extent than OH-ion. Ca-ion enters the soln. instead of Ba; therefore this is a case of substitution adsorption. This satisfies Freundlich's equation $x/m = \alpha C^b$. Comparison of adsorption of Ba from solns. of $Ba(OH)_2$, $BaCl_2$ and $BaCl_2 + HCl$ showed a decrease of adsorbed Ba with increase of H-ion concn. In agreement with this, Th ion is adsorbed less than Al ion and less than Ba ion, from solns. of equal concn. as its hydrolysis results in a higher concn. of H ions. J. G. TOLPIN

Adsorption of chlorides of aluminum and thorium by clays and kaolins. N. A. HELD. *J. Russ. Phys.-Chem Soc* 62, 1553-70 (1930).—H. studied the character of adsorption of these salts by 5 kinds of Russian clays and made a no. of electrometric titrations of the ions of $AlCl_3$ and $ThCl_4$. A no. of tables and curves show that Cl^- is adsorbed in negligible quantities; while Al and Th are adsorbed in different proportions by different clays (from 1 to 10 times). The general character of adsorption shows that with increased concn. of the cation the quantity adsorbed approaches a limit and seemingly satisfies the isothermal adsorption equation of Freundlich. The cations Ca^{++} and Mg^{++} substitute Al^{+++} and Th^{++++} in the soln., i.e., substitution

adsorption takes place. It is suggested that the clays can be characterized by the amts. of adsorbed aluminosilicates. Displacement of hydrolytic equil. results in acidity of the soln. The hydrolysis is expressed by: $\text{AlCl}_3 + 2\text{H}_2\text{O} \rightleftharpoons \text{Al}(\text{OH})_2\text{Cl} + 2\text{HCl}$
 $\text{ThCl}_4 + 3\text{H}_2\text{O} \rightleftharpoons \text{Th}(\text{OH})_3\text{Cl} + 3\text{HCl}$. Detd. constts. for the hydrolysis of AlCl_3 at 20° range from 1.1×10^{-12} to 1.95×10^{-12} ; at 25° from 2.5×10^{-12} to 3.6×10^{-12} , for ThCl_4 at 20° 5.5×10^{-12} to 5.9×10^{-12} . J. G. TOURNE

Adsorption of easily soluble and difficultly soluble electrolytes on precipitates with large surface. L. LERNER. *Z. anorg. Chem.* 43, 575-7 (1930).—In adsorption phenomena on pptts. with large surface generally 2 processes are involved. The process first ("momentary" adsorption) is an ion adsorption that depends on the charge on the surface and the valence of the adsorbed ion. The second process consists of a building up of mols. and re-soln. of the resulting mols. at the surface layer. In this process the velocity of the decrease of surface and the mols. of the adsorption compd. are important factors. Curves are given showing the adsorption of Pb (Th B) and actinium on AgCl , AgBr and AgI as a function of time. If the systems are brought together immediately on prepn. adsorption increases with time otherwise it steadily decreases. With AgCl and AgBr the adsorption increases quickly to a max. and then gradually decreases. R. I. REYN

Problems of the boundary state. Wm. HARDY. *Trans. Roy. Soc. (London)* A230, 1-37 (1931).—More than 100 substances, chiefly long-chain compds. including normal paraffins, acids and alcs., and carbonols, were used as lubricants in expts. to det.: (1) the static friction between solid surfaces, either clean or separately coated with solid lubricant, (2) the static friction of clean surfaces joined by fluid lubricant, (3) the tangential force needed to produce slip in a solid joint between a cylinder and a plate, (4) the normal force to break a fluid joint between a cylinder and a plate, (5) the normal force needed to break the same joint when solid. The thin films of lubricant between the solid surfaces may be regarded as a fourth state of matter since the energy of these films is shared so overwhelmingly with the enveloping phases. For each load there is a value h , sepp. cylinder and plate, at which the capillary pressure is equal to the loading. This value is const. for all fluids including clean air and for different solids when the surfaces are clean. This capillary pressure is called the Leslie pressure, it may be accounted for (1) by direct attraction of relatively long range as contemplated in the Laplacian theory of short range forces, or (2) by the spreading from mol. to mol. of a state of strain. Mols. that are oriented and strained and in turn can orient and strain mols. beyond them are in a polarized state and the spreading of polarization from mol. to mol. is termed diachysis. The Laplacian theory seems to account most satisfactorily for the Leslie pressure in clean air, but in liquids the most promising view of the boundary state is that it is due to the formation of chains of polarized mols. stretching through the lubricant from one boundary surface to the other. Each chain has little strength in shear and great strength in tension, but both shear and tensile strength decrease as the length of the chain increases. C. C. KRISS

Anomalous velocity distribution in thin lubricant films. H. UHLMANN. *Kolloid Z.* 54, 220-6 (1931).—The Hagen-Poiseuille equation assumes optically thin layers and 0 velocity at the wall layer. These assumptions are not true in lubrication, where processes at the wall may not be neglected. An attempt is made to derive a law by math. summation of the forces acting between the wall and the various layers. Dynamic viscosity: the true measure of the value of an oil for lubrication is the sum of the true viscosity and the adhesiveness, the specific combination between oil and metal. At present it is only possible to get an indication of the adhesiveness from the temp.-fluidity curves of the lubricant. The fluidity of oils with symmetrical molecules varies as the square of the temp.; for those with unsymmetrical molecules it varies as the temp. A flat temp.-fluidity curve signifies orientation and "adhesiveness" (anomalous velocity distribution). The viscosities of a Baku spindle oil, a Pennsylvania oil, a *Rheumonia fraction* 310-350°, *Shell 1 distol*, *caster oil* and *rapeseed oil* were detd. with a microviscometer from 30° to 200°. The results may be expressed by a semi-empirical rule $\eta = (C M^2)/(t - t^*)^n$ where C is a characteristic const. for a homologous series, M is the mol. wt., t^* is a reference temp. and n a value from 1 to 2 depending on the elec. symmetry of the mol. ARTHUR FLEISCHER

Adhesive forces on surface films. T. M. LOWRY. *Nature* 127, 165 (1931).—L. explains Adam's contention (*C. A.* 25, 1422) that the mols. of a fatty acid on the surface of water may be held together by mutual attraction of the heads on the basis of the conception of the coordination of H. MALCOLM DOLE

The film state of matter. A. V. BLOW. *Kolloid Z.* 54, 210-20 (1931), cf. *C. A.* 23, 1290, 2049, 4087, 5536.—Films stand between dispersoids with large unconnected sur-

faces and solid bodies with small but connected surfaces. The types of film are the uni-molecular (Langmuir, etc.), Hardy films which are thicker than the active range of molecular forces, and zone films. Films may be formed by physical and by chemical processes. The kinetics of film formation by evaporation are reviewed. A. F.

Elliptical polarization by reflection at the surface of liquids: application to the study of monomolecular surface films. CH. BOUCHER *Ann. phys.* [10], 15, 5-130 (1931).—A study of elliptical polarization by reflection at Brewsterian incidence from the surface of water, salt solns., pure org. liquids and solns. of fatty acids. The ellipticity (K) is shown to be pos. in all these cases. K for salt solns. is identical with that for pure H_2O (0.00040), confirming the presence of a Gibbs' layer of adsorbed H_2O at the boundary surface. Surface films of fatty acid solns. have a thickness calcd. from K and Drude's theory of the order of mol. dimensions and particularly of approx. the same magnitude as d , the mean distance between the mols. as detd. from x-ray studies, indicating their monomol. nature. In the case of monomol. fatty acid films K consists of 2 terms, K_1 pos. and relating to the H_2O or other liquid supporting the film and K_2 neg., due to the film and relating to the length of the mols. composing the film. In the case of pure liquids K is given as a first approximation by a relationship of the form, $K = d \times F(\lambda, T, \dots)$. A. P. SACIS

Certain mechanical properties of surface tension. V. DUCESCU *Bull. soc. chim. Roum.* 5, 1069-76 (1930), cf. *C. A.* 25, 1136.—The formation of adsorption films in the interface liquid (H_2O)-gas (H_2O vapor) is accompanied, for certain substances, by the production of a marked quantity of mech. energy, capable among other things of producing a regular and continuous movement of a small rotor with rotating blades on the water. This mech. factor, along with the great velocity with which mol. dispersion of almost bidimensional form takes place, should be considered in physiol. processes in which surface-tension energies predominate, e.g., muscle contraction. Diagrams and description of the app. are given. PETER MASUTCI

A microscopic method. E. H. SYNGE. *Phil. Mag.* 11, 65-80 (1931), cf. *C. A.* 22, 4048.—The use of a colloidal Au particle at the focus of a slightly elliptical mirror is proposed as the source of a spot of intense light, very small in comparison with the wave length. A particle 10^{-4} cm. in diam. will scatter sufficient light to operate a photoelec. cell. This method would increase the resolution 10 to 20 times beyond that of the ordinary microscope. An instrument using this principle is named an *hypermicroscope*. ARTHUR FLEISCHER

Measurements of suspensions and deposits. PIERRE MERCIER. *J. phys. Radium* [7], 1, 292-305 (1930).—A float is attached by means of a hair to a lever. It is immersed in homogeneous or heterogeneous suspensions. While the suspensions settle, the movement of the lever is optically recorded. The results are presented in curves upward pressure on the bob vs. time, and pptn. velocity vs. concn. A. L. H.

Further progress in ultrafiltration. ERWIN KRATZ. *Chem.-Ztg.* 55, 257-9 (1931).

E. H.

Form and structure of colloidal particles. A. BOUTARIC. *Rev. g^e sci.* 42, 11-9 (1931).—A descriptive paper summarizing the modern conception of colloidal particles.

A. LLOYD TAYLOR

The explanation of Stokes' rule. FRANZ URBACH. *Sitzb. Akad. Wiss. Wien, Abt. IIa*, 139, 473-81 (1930).—Theoretical. The relationship between Stokes' rule and the sec. law of thermodynamics is discussed, with the use of an idealized cyclical process. It is shown that deviations are possible through the effect of the heat energy. It is necessary to assume that the emission mechanism is independent of the heat energy. E. J. R.

Temperature and the stability of colloidal solutions. S. I. D'YACHKOVSKI. *Kolloid.-Z.* 54, 278-84 (1931).—The absorption spectra of well-dialyzed hemoglobin and albumin sols were detd. from 20° to 100° at 10° intervals. All solns. after coagulation showed a Tyndall cone. Surface tensions were detd. for a no. of sols from 10° to 70°. The curve of $\Delta\sigma/\Delta t$ against the temp. shows a max., which is very decided for Fe_2O_3 sols at 30-40°. A large no. of sols were subjected to freezing at temps. from -20° to -182° and the coagulation was noted. The surface tension and cond. of the thawed sols were detd. The behavior of the sol upon freezing is a good indication of its stability. Uranium and tungstic acid sols were fractionated by use of successive lower temps. of freezing.

ARTHUR FLEISCHER

Contributions to colloidal syntheses with well-crystallized organic compounds. NADINE VON VERNERN. *Kolloid.-Z.* 54, 296-306 (1931); cf. *C. A.* 24, 2357, 2657, 3415, 3506.—Colloidal solns. of papaverine, camphor, salol, benzophenone, anthracene, an-

of α -phenanthrene in aq. sugar sols are described. Freshly prepd. solol
 of α -phenanthrene sols have spherical liquid particles. In the coarse particles the
 of α -phenanthrene can be detected. ARTHUR FLEISCHER

Photometric measurements of concentration and dispersion in colloidal solutions.
 III. J. E. STEN L. HILL. *Poland J.* 54, 150-6 (1930), cf. *C. A.* 25, 870.—The use of
 of α -phenanthrene was studied by means of mastix sols. If the incident radiation is
 of α -phenanthrene weakening const. a max. appears. The curves at different wave
 of α -phenanthrene are nearly the same. Slight deviations are due to errors in the prism factor and
 of α -phenanthrene effect. K. L. HILL's rule with const. dispersion holds up to the value of the
 of α -phenanthrene const. of 1. The usefulness of nephelometry ceases where light-weakening
 measurements begin. The red part of the spectrum is most useful for white sols.
 of α -phenanthrene is very sensitive to changes in the dispersity. ARTHUR FLEISCHER

Investigations of colloids, prepared by the method of condensation of molecular
 rays. I. Organosols of the alkali metals. ADOLPH I. RABINOVICH, V. A. KARGIN AND
 J. I. LITMAN. *Poland J.* 54, 208-65 (1931).—A new app. is described for the prepn.
 of more stable sols of the alkali metals in org. solvents. K. sols in ether can be prepd.
 of α -phenanthrene in 1 to 2 to 3 mos. Cataphoretic measurements indicate that the particles
 are neg. charged. K. increases the stability and also the cataphoretic velocity. Light
 causes the pptn. of metal from the pure sol, and a blue gel from the protected sol. K.
 is in benzene, xylene and tripropylamine are very unstable. K. sols in ethylamine are
 stable for 3-6 hrs. and are characterized by the presence of K in true sol. A. P.

Dielectric constant and structure of thixotropic sols. S. S. KESTER. *J. Phys.
 Chem.* 35, 815-29 (1931).—The dielec. const. of thixotropic Al_2O_3 , Fe_2O_3 , and gelatin
 sols were measured with 32.7-cm. waves by the second Drude method. In all cases the
 dielec. const. decreased with increasing concn. of electrolyte (K_2SO_4). The absorption
 of the elec. waves was practically the same for the sols as for pure water. Pure K_2SO_4
 solns., which were measured for comparison, showed a negligible decrease of dielec.
 const. and a definite increase of absorption. The reversible sol-gel transition of a colloid
 is not accompanied by discontinuous changes in its dielec. properties. L. ONSAGER

Specific inductive capacity and molecular weights of colloids. NINA MARINOVSKA.
Compt. rend. 192, 625-8 (1931).—An equation is derived from the Debye relation which
 permits detg. the mol. wt. of dissolved proteins from the dielec. dispersion of the system
 and the wave length. The mol. wt. obtained by this method corresponds to the smallest
 particle of dissolved substance even if it is in the presence of more complex molecules that
 possess a high degree of assocn. The mol. wt. of gelatin is 11,200 by this method.

J. R. ADAMS

Dispersoidological study of silver salts in aqueous ethyl alcohol, acetone and propi-
 one solutions. I. SMAGIUS ISHII. *Bull. Chem. Soc. Japan* 6, 53-60 (1931).—The
 complex soly. (soly. in the form of complex ions) of pptd. AgI in aq. EtOH solns. of KI
 and in aq. $(CH_3)_2CO$ solns. of KI was detd. The dispersion of the AgI in these sols
 was also studied. The soly. is greater the higher the proportion of EtOH or $(CH_3)_2CO$
 in the solvent. The greatest amt. of dispersion and the widest range of KI concn.
 through which there is dispersion is found at intermediate proportions of EtOH or
 $(CH_3)_2CO$. P. L. BROWN

Investigations on the aging of aqueous ferric chloride solutions. A. LOTTERMOSER
 AND ERICH LESCHKE. *Kolloidchem. Behefte* 32, 157-68 (1931).—The changes in cond. of
 dil. solns. (0.00259 to 0.00000221 M) of $FeCl_3$ in cond. water were studied. The cond.
 does not attain a const. value at the end of 2500 hrs. The cond. of the more concd.
 solns. at that time is increasing slowly, while that of the more dil. has passed through a
 max. and is slowly decreasing. A rise in temp. increases the rate at which these changes
 go on, but a temp.-reversible equil. was not observed. The same is true for an aging
 Fe_2O_3 sol. Colorimetric detns. show that light is more strongly absorbed by $FeCl_3$
 solns. as the wave length becomes shorter and the time of aging increases. J. W. P.

Studies in electroendosmosis. IV. The electroendosmosis of some organic
 liquids against a glass surface. FRED FAIRBROTHER AND MARK BALZIN. *J. Chem.
 Soc.* 1931, 389-403.—Measurements were made of the electroendosmosis of fourteen
 pure org. liquids and water through a diaphragm of sintered fena-glass powder. C_6H_6
 and CCl_4 show no measurable effect, Et_2O is charged neg. toward the glass and all the
 other liquids are pos. charged. Surface cond. was found in the glass diaphragm in the
 case of each liquid that showed endosmosis. Up to 150 volts the relation between the
 voltage and the velocity of electroendosmosis was linear, but at higher voltages the ve-
 locity increased more rapidly than the voltage, which may be caused by the increased
 strom polarization experienced by the surface layer of the liquid in the diaphragm. In
 the cases where the dipole moment of the liquid, μ , was known, the function $(D/\Delta f)^{1/2}$

), where D is the d , M mol wt and η viscosity, gave a straight line when plotted vs the velocity of electroendosmosis per volt.

Studies on electrokinetic potentials. VIII. Ion antagonism. HENRY B. BULL, ROSS AIREY GORTNER. *J Phys Chem* 35, 700-21(1931), cf *C A* 25, 1721—electrokinetic potential for mixts of KCl and $NaCl$ is merely a component of the individual potentials, no ion antagonism is evident. Similarly the results clearly state the absence of ion antagonism between $MgCl_2$ and $CaCl_2$, as influencing the electrokinetic potential.

The influence of surface active compounds and electrolytes, with consideration of types of ions, on the electrophoretic velocity of lyophobic sols. SIAR HONG WHANG. *Kolloidchem. Beihefte* 32, 169-96(1931)—The electrophoretic velocity (cataphoresis) of various lyophobic sols (As_2S_3 , quartz suspension and Fe_2O_3), to which various amounts of different electrolytes or surface active compounds (alcs, amines and fatty acids) had been added, was determined by the moving boundary method. No well-defined relation was observed between surface activity and electrophoretic velocity. The addition of quartz suspension of an alc up to butyl or a fatty acid caused a decrease in electrophoretic velocity, while amyl and hexyl alcs as well as various organic bases, both surface active and not, had the opposite effect. The expts with electrolytes show that the electrophoretic velocity of charged particles decreases with increasing mobility of the oppositely charged ions but increases with increasing mobility of the ions of like charge.

Effect of alkali salts on the cataphoresis and precipitation of colloidal gold. A contribution to the study of the Hofmeister series. ANNA LAGEMANN. *Kolloidchem. Beihefte* 32, 212-48(1931)—The effect of various alkali halides on the cataphoresis of colloidal Au was determined by a carefully worked out ultramicroscopical method. The results show that the charge on the Au particle was decreased in the order $Li < Na < K < Rb < Cs$.

The negative ions hindered this discharging action in the order $Br < Cl < F$. The charge on the particles did not completely vanish but approached an asymptote as the concentration was increased from 2×10^{-4} to $50 \times 10^{-4} N$. This asymptotic charge was increased as the alkali metal ion increased in atomic weight. The change in charge with varying concentration can be represented by an adsorption curve. The coagulating effect of the various salts was determined by observing color changes. Those ions most effective in discharging Au sol were also most efficient in causing coagulation. No quantitative relation was found, however, between the two effects. The varying effects observed with different ions can be explained by differences in their ionic radius, mobility and degree of hydration. Comparison with similar effects observed with lyophilic colloids shows that the same Hofmeister series holds for both types of colloids.

Contribution to the study of the action of electrolytes added to a colloidal solution in proportions below those necessary to cause coagulation. G. ROSSI and A. MARESCA. *Gazz. chim. ital.* 60, 693-6(1930)—Earlier expts showed (cf *C A* 21, 2411) that electrolytes that normally are stabilizers (e.g., H_2SO_4 and Na_2SO_4 for colloidal S) function as stabilizers only within certain limits of concentration, and at higher concentrations become coagulants of the same sols. With a given colloidal soln, a given electrolyte requires a certain quantity necessary to cause coagulation, and another electrolyte for coagulation, as for the 1st electrolyte, with respect to the 2nd, an optimum concentration that brings about the maximum stabilization. Above this, but with quantities below those necessary for coagulation, the electrolyte either does not bring about any change or brings about a slight stabilization. This was demonstrated by expts. on colloidal As_2S_3 sols, and for the sake of simplicity with HCl both as the electrolyte, added in quantity below that necessary for coagulation, and also as coagulant. It was found that when added to a colloidal As_2S_3 soln in quantity below that necessary for coagulation, HCl can either stabilize or destabilize the soln toward the coagulating action of the HCl itself, according to the concentration of the HCl . At the concentration of HCl which stabilizes the soln, there is an increase in the degree of dispersion, whereas the dispersion is diminished by the addition of a quantity of HCl which, though not reaching the minimum necessary for visible coagulation, destabilizes the colloidal soln toward the coagulating action of the HCl itself. The colloidal As_2S_3 soln was prepared by passing H_2S for a long time into an aq. suspension of As_2O_3 , diluting and removing excess H_2S with a current of inert gas, and dialyzing. The results of dispersion and the coagulation phenomena after the various treatments were determined by the ultramicroscope.

Coagulation kinetics of suspensions. I. K. JABLONSKY. *Kolloid-Z.* 54, 164-9(1931)—Modification of Smoluchowski's assumption by assuming that in coagulation the rate of reaction is established at which all particles are charged, leads to the equation

time t and z the no. of charged particles. When $z = \text{zero}$ the equation is the same as S. 5. The equation holds well for the data of Westgren and Reistötter on the coagulation with $S = 2$. Spectrophotometric measurements of the rate of coagulation of FeCl_3 , Ag_2S , SnS and AgCl sols by KCl at 18° agreed well with the formula.

ARTHUR FLEISCHER

Viscometric investigations of hemoglobin coagulation. S. L. PERKO. *Kolloid-Z.* 54, 170-5 (1931); cf. *C. A.* 24, 137. —The effect of $\text{C}_2\text{H}_5\text{OH}$ and $\text{C}_4\text{H}_9\text{OH}$ on coagulation by KCl , CaCl_2 and FeCl_3 was studied by the detn. of viscosity at 17° . In the presence of KCl the viscosity remains const. or decreases; in the presence of $\text{C}_2\text{H}_5\text{OH}$ the viscosity remains const. or increases. The valence effect of the cation in the presence of alc. is only slightly noticeable. At high concn. of CaCl_2 with $\text{C}_2\text{H}_5\text{OH}$ stabilization of the system was observed. Stabilization occurs with FeCl_3 and $\text{C}_4\text{H}_9\text{OH}$ from concns. of 0.1 % to 2 % above which coagulation makes viscosity measurements impossible. Also in *J. Russ Phys. Chem.* 62, 2573 (1928).

ARTHUR FLEISCHER

Pectography—the study of the progressive drying of solutions. PAUL BARY. *J. Am. Phys.* 25, 1-13 (1931); cf. *C. A.* 25, 1676. —The structure of deposits left on a glass slide placed vertically in a beaker contg. an aprotic colloidal soln. was studied. In general hydrophobic sols deposit a series of ridges resembling Liesegang rings while the residues from hydrophilic sols show no structure. Most colloids belong strictly to neither class but represent intermediate systems and may approach either autopeptization or coagulation on aging. Fe_2O_3 residues consist of two phases, one composed of doubly refracting filaments arranged parallel to each other and imbedded in an amorphous phase.

J. G. McNALLY

Periodicity and its fundamental principles. MATTHEW CORNAROW. *Kolloid-Z.* 54, 257-65 (1931). —Periodic formations are obtained when a CO_2 stream is solidified by cooling, in the sublimation of $(\text{NH}_4)_2\text{CO}_3$, and in burning Mg in oxygen. Yellow P did not show a rhythmic structure either on oxidation or on sublimation in a CO_2 stream. Periodic structures are the outer result of a periodicity and are developed only in a system where there is harmony between wave frequency and amplitude and the natural frequency of the system.

ARTHUR FLEISCHER

The formation of Liesegang rings. (Miss) S. ROY. *Kolloid-Z.* 54, 190-3 (1931). —The following new rings were formed: HgI_2 in V_2O_5 gel, HgI_2 and Ag_2CrO_4 in CeO_2 , TlI and HgI_2 in ZnAsO_4 , TlI in MnAsO_4 , BaCrO_4 , TiCrO_4 , AgI in SiO_2 , TlI and AgI in starch, and TlI and CuI in agar. Liesegang rings are more marked when formed in the presence of light than in the dark, because of the accelerating effect on coagulation by light.

ARTHUR FLEISCHER

Microscopic and cinephotographic study of Liesegang rings. SUZANNE VEIL and L. BILL. *Compt. rend.* 192, 282-4 (1931). —The rings formed by the action of a drop of AgNO_3 on dichromate-impregnated gelatin formed concentrically with pointed protuberances interrupting their continuity. As successive rings were formed the protuberances receded but the rings were left incomplete at these points. The velocity of formation of the rings decreases and the intercepting protuberances disappear as the contact line between the solns. recedes toward the center.

A. LLOYD TAYLOR

Practical method for determining vapor-tension diagrams. G. CALINGAERT. *Chimie & industrie* 25, 307-8 (1931). —A brief description and explanation are given of a method recently introduced in America; cf. *Davis, C. A.* 24, 4967.

A. P.-C.

Calcium acetate gels. H. C. G. SMITH. *Kolloid-Z.* 54, 181-90 (1931); cf. *C. A.* 23, 4393. —Further expts. were carried out on the formation of the gels with the use of pyridine, Me Et ketone , etc. The addn. of org. liquids immiscible with water shows no visible action in the formation of gels. $\text{Mg}(\text{ClO}_4)_2$ produces a cryst. ppt.; because of alcoholate formation CaCl_2 has no visible effect, Na and Mg acetates form gels quickly. Photomicrographs show that most of the gels have a honeycomb structure with $\text{Ca}(\text{AcO})_2$ as the disperse phase. The gels formed by acetone in the presence of $\text{C}_2\text{H}_5\text{OH}$, CS_2 show a phase reversal. The $\text{Ca}(\text{AcO})_2$ disperse phase on aging increases size and is accompanied by opalescence. The crystals finally formed may be the anhyd. or monohydrated $\text{Ca}(\text{AcO})_2$. III. *Prod.* 284-9 (1931). —The viscosities of $\text{C}_2\text{H}_5\text{OH}-\text{H}_2\text{O}$ and $\text{C}_2\text{H}_5\text{OH}-\text{H}_2\text{O}-\text{Ca}(\text{AcO})_2$ solns. were detd. from 15° to 40° in an Ostwald viscometer. For the alc. water solns. the data agree with the literature, a max. occurring in each of the viscosity temp. curves. The viscosity of the Ca acetate solns. increases with the Ca concn. The change from soln. to gel is very sudden, there is no gradual transition in viscosity. The general decrease in viscosity upon addition of alc. shows that the effect is mainly a dehydration of the Ca acetate. Other factors in gel formation are the lowering of the surface tension and the decreased dissociation of Ca acetate in the presence of alc.

ARTHUR FLEISCHER

Investigations on gum arabic with the aid of colloidal silver and gold. J. VOIGT. *Kolloid-Z* 54, 307-10(1931); cf. *C. A.* 23, 3152—Very pure gum arabic soln. does not reduce Ag_2O or AuCl_3 solns. Preliminary illumination of the gum arabic soln. decreases its protective action on Ag and Au sols. Arabic acid behaves similarly. Ag sols cannot be prepd. by radiation when gum arabic or arabic acid is used as the protective colloid. The nature of the changes in the gum arabic is unknown. A. F.

The effect of salts on the ionization of gelatin. KENNETH V. THIMANN. *J. Gen. Physiol.* 14, 215-22(1930)—The addn. of NaCl , CuCl_2 , and CaCl_2 to gelatin solns. is shown from the Donnan membrane equl. to increase the ionization of gelatin. The effect is due to the formation of complex ions. The evidence supports the amphoteric ion structure of protein mols. rather than the classical theory of structure. C. H. R.

Diffusion in gelatin gels. WALTER STILES AND GILBERT S. ADAIR. *J. Am. Chem. Soc.* 53, 619-20(1931)—The results of S and A (*C. A.* 16, 800) on the decrease in the rate of diffusion of electrolytes in gelatin gels as the concn. of the latter is increased agree essentially with the recent results of Friedman and Kramer (*C. A.* 24, 2934).

The index of refraction of gelatin solutions and the supposed hydration of the dispersed particles. G. ROSSI AND A. MARESCOTTI. *Gazz. chim. ital.* 61, 14-26(1931).—A review of the literature on colloidal solns. of gelatin (to which numerous references are given) suggested that the expts. of Walpole (cf. *C. A.* 8, 859) on the relation between the n value and the concn. of gelatin be carried out with the greatest precision possible. Accordingly sterile aq. solns. of pure gelatin of various concns. at 18° were measured with an immersion refractometer, the technic of which is described. The results are tabulated and discussed in detail and lead to certain general conclusions: (1) Gelatin dissolved in water increases the n value of the latter, but this increase is not a linear function of the concn. of the gelatin. (2) This divergence from the linear relation is a confirmation of the hypothesis of R. (cf. *C. A.* 18, 3510) and of Kruyt (cf. *C. A.* 23, 3388), Kruyt and Tendeloo (cf. *C. A.* 19, 919, 3192) on the solvation of the particles of lyophilic colloids. (3) The differences in the n value of water as a result of the dissolved gelatin found at different temps. probably depend upon variations in the degree of hydration with change of temp. which are in conformity with the general behavior of lyophilic colloids. (4) Variations in the n value of gelatin solns. by addn. of salts (Na_2SO_4 , NaI , NaBr and Na citrate were tested) are not strictly additive. The salts behave differently, in that both increases and diminutions in the viscosity of the gelatin solns. were found, and the variations caused by these salts are such that no explanation can be offered at present. C. C. DAVIS

Structure of celluloid and nitrocellulose and the gelatinizing medium of nitrocellulose as a swelling medium. II. X-ray spectrograms of camphor celluloid with varying camphor content. J. R. KATZ, J. C. DERKSEY, C. A. KRAMERS, KURT HESS AND CARL TROGUS. *Z. physik. Chem., Abt. A*, 151, 145-62(1930), cf. *C. A.* 24, 5201—Camphor in celluloid acts as a lubricating agent between the micelles of nitrocellulose. In addn. to this, it forms with the latter, compds. sufficiently definite to show breaking up of the diffraction rings into segments, if the celluloid film is stretched. Typical spectrograms are reproduced and discussed. III. Celluloids with other cyclic ketones as jelling media. J. R. KATZ, J. C. DERKSEY, KURT HESS AND CARL TROGUS. *Ibid.* 163-71 (1930)—Fenchone, menthone, carvone, cyclopentanone, cyclohexanone and cyclohexanol also form these addn. compds. with nitrocellulose. A tabular correlation is given of the ring diams. and the interplanar distance d , for various percentage mixts. of these compds. with nitrocellulose. IV. Acid amides and esters as jelling media. *Ibid.* 172-89(1930)—These compds. likewise form addn. compds. with nitrocellulose. Since ketones, amides and esters have only the carbonyl group in common, it is concluded that it is by means of the latent valences of this group that the addn. takes place. GREGG M. EVANS

Thermoelastic effect in cellulose ester films. J. G. McNALLY AND S. E. SHEPARD. *J. Phys. Chem.* 35, 100-14(1931)—At room temp. films of cellulose nitrate or acetate cool when stretched to small elongations but give off heat when elongated beyond their yield point. With strongly biaxial films the secondary heating effect is absent. Under proper stress and temp. conditions cellulose nitrate films contract when heated and the coeff. of thermal expansion depends on the previous mech. treatment of the material, the temp. range, the rate of heating and the stress on the film. A value of 6×10^{-3} is given for the coeff. of linear expansion at low stresses between 15° and 45°. The results are related to micellar structure. J. G. McNALLY

Solubility of nitrogen in water at high pressures and temperatures. JOHN B. GOODMAN AND NORMAN W. KRASE. *Ind. Eng. Chem.* 23, 401-4(1931)—A high-pressure

Changes of the osmotic water attraction of a system of phases. I. F. A. II. SCHRIFMACKERS *Proc Acad Sci Amsterdam* 33, 694-701 (1930), cf following abstr. —The problem of the general osmotic system $E E'$ is discussed, where both E and E' may represent any no. of phases in equl. Expressions are set up for the thermodynamic potential of the whole system and the following results are deduced: (I) When the pressures or vols remain const., on both sides of the membrane, the H_2O diffuses toward that side on which the osmotic water attraction ($O W A$) is the greater and no water diffuses when the $O W A$ is equal on both sides. (II) An increase in the pressure of an equl. (i.e., E or E') for which the vol. increases with absorption of water results in a decrease in its $O W A$, and a decrease in vol. produces a decrease in the $O W A$. (III) When an equl. takes in water at const. pressure or at const. vol., its $O W A$ will decrease unless all phases keep the same composition, and in general for a given quantity of water the effect at const. pressure is greater than that at const. vol. From these it follows that (IV) increase in pressure or decrease in vol. of a liquid or vapor decreases its $O W A$ and (V) when a liquid or a vapor takes in water at const. pressure or at const. vol., the $O W A$ is decreased and for a given quantity of water the effect at const. vol. is greater than at const. pressure. The paper is highly mathematical. W. W. S. SIFER

Equilibria in osmotic systems, in which forces act. II. Osmotic systems with an active membrane, permeable for several substances. I. A. II. SCHRIFMACKERS *Proc Acad Sci Amsterdam* 33, 1133-4 (1930) cf preceding abstr. and *C A* 24, 4417. —The general problem of osmosis deals with an active membrane permeable to more than one substance. An abstract mathematical treatment of the thermodynamics of this case is presented and a no. of conclusions are deduced, for which the original paper should be consulted. W. W. SIFER

New micro method for the ebullioscopic determination of molecular weights. JAMES H. C. SMITH AND HAROLD W. MUNER *Mikrochemie* 3, 117-22 (1931). —(In English.) An app. and its use are described which is based upon the larger app. of Menzies and Wright (*C A* 16, 2340). Satisfactory results can be obtained by using as little as 3 cc. of solvent and 5-25 mg. of solute. It does not require a supply of Pt tetrahydrazide, as does the Beckmann micro method, substitutes a differential thermometer for the much more expensive and more difficultly obtainable Beckmann thermometer and the readings are easily obtained because the observed change is greater. The app. consists of a boiling tube, the differential thermometer and a Cottrell pump. Directions are given for carrying out a detn. and the results obtained shown for urethan, *p*-toluidine, vanillin, diphenylamine, naphthalene and triphenylmethane. W. T. H.

The surface tensions of aqueous solutions of *p*-toluidine. R. C. BROWN *Phil Mag* 11, 150-60 (1931). —The surface tension of *p*-toluidine solns. with 0-0.6 g. per 100 cc. of soln. was detd. at $\pm 0.05^\circ$ by the ring Ierguson, capillary rise and pull-on sphere methods with an accuracy of 0.5%. The results are in fairly close agreement with Gans and Harkins. The results by the Ierguson method (*C A* 16, 2136) do not agree with those obtained by the same method by I. Edwards (*C A* 19, 3185). A. I.

The viscosity and density of rubidium nitrate solutions. HAROLD G. SMITH, JOHN H. WOLFFENDEN AND HAROLD HARTLEY *J Chem Soc* 1931, 403-9. —The relative viscosities (by the method of Merton, *C A* 5, 1221) and the relative d_s (by the method of Hartley and Barrett, *C A* 5, 2096) of 0.092-1.788 *N* aq. $RbNO_3$ solns. were measured at 18° and 25.01° . The viscosity concn. curves for $RbNO_3$ solns. at both temps. are similar to those of Griegsen for KNO_3 and of Merton for $CsNO_3$. For concns. of $RbNO_3$ solns. below 0.2 *N*, the relation between relative viscosity and concn. is represented by the Jones-Dole equation (*C A* 23, 5375). The graph of $(\phi - 1)/\sqrt{c}$ against \sqrt{c} for $RbNO_3$ solns. at 18° and 25° shows that both the intercept A and the temp. coeff. are neg. The "mol. soln. vol." is calcd. for $RbNO_3$ at various concns. of solns. from the d measurements and is shown to fall in the normal periodic order of the alkali nitrates. J. BALOZIAN

The behavior of electrolytes in mixed solvents. II. Effect of lithium chloride on the activities of water and alcohol in mixed solutions. R. SHAW AND J. A. V. BUTLER *Proc Roy. Soc. (London)* A129, 519-35 (1930). —An air-bubbling method, in which the vol. of air required is less than 2 l., was devised for the detn. of the partial pressures of H_2O and alc. in mixed solns. Measurements were made at 25° of the partial vapor pressures of H_2O -alc. solns. contg. from 0 to 100% alc. and 0.5, 1.0, or 4.0 mol. % While increasing concns. of $LiCl$ cause a decrease in the partial pressure of H_2O in all solns., its effect on the alc. varies with the composition of the solvent. In solns. contg. a large proportion of alc., its partial pressure falls steadily as the salt concn. increases, and in solns. contg. 6.4 mols. % of alc., it rises steadily, while in intermediate solns. it first falls and then rises. The fractional lowering of the vapor pressure of water is

always greater than that of alc. An approx linear relationship is established between the relative activities of both water and alc. and the mol fraction of alc. in the solvent, at const LiCl concn. It is possible to distinguish approx the effects of the soln. of ions by alc. mols and the salting out of alc. by the interaction of the ions and water.

A. L. HEYNE

The fictive volumes of sodium sulfate in aqueous solutions of sulfuric acid and of iodine in an aqueous solution of potassium iodide. R. E. GIBSON. *J. Phys. Chem.* 35, 690-8(1931).—The sp. vols. of 2 series of solns. of Na_2SO_4 were detd. in the first series approx 5% aq. H_2SO_4 and in the second approx 10% H_2SO_4 , being the solvent. The fictive vols. of Na_2SO_4 in these solns were calcd. Peculiarities of the vol.-conc. curves indicate the formation of NaHSO_4 in soln. Measurements of the sp. vols. of I in a 49% soln. of KI showed that in solns. contg. 10 to 50% of I, the fictive vol. of I is independent of the concn. of I, and close to the most probable estimate of the sp. vol. of liquid I at 25°.

L. OVSAGER

The variation of the color of solutions of cobalt chloride. E. TOROJESCU. *Compt. rend.* 192, 290-2(1931).—Solns. of CoCl_2 7.17 g./l. change from blue to rose or rose violet at the temp. indicated for each solvent: MeOH, 3°, EtOH, -18°, PrOH, -32°, BuOH, -45°, AmOH, -65°. For the same solvents the dielec. consts. are 34.5, 32.5, 28.5, (37.5), 25.0. A relation probably exists between the decreasing dielec. consts. and a variation in degree of polymerization which accounts for the color change.

A. LLOYD TAYLOR

The hydrolysis of zinc sulfate solutions, solubility product of hydrous zinc oxide and the composition of the latter precipitated from zinc sulfate solutions. I. M. KOLZNOFF AND TOSUO KAMEDA. *J. Am. Chem. Soc.* 53, 832-42(1931).—By using very pure materials it was found that the 2nd ionization const. of $\text{Zn}(\text{OH})_2$ at 25° is $4.4 (=0.4) \times 10^{-14}$. The soly. of hydrous ZnO pptd. from the sulfate is 1×10^{-4} at 25°. On pptn. of the hydrous oxide from sulfate soln. with NaOH the basic salt $\text{ZnO} \cdot \text{ZnSO}_4$ is formed. After heating to 80° for 10 days the ppt. becomes more basic and the supernatant liquid more acid.

CHANNING WILSON

The transport numbers of hydrogen chloride in ethyl alcohol. J. W. WOOLCOCK, HAROLD HARTLEY AND O. L. HUGHES. *Phil. Mag.* 11, 222-5(1931), cf. *C. A.* 22, 3084.—The e in f of the cell $\text{Ag} \cdot \text{AgCl} \mid \text{HCl } m_1 \mid \text{HCl } m_2 \mid \text{AgCl} \cdot \text{Ag}$ was measured at 25°. The most probable value for the transport no. of H ion at infinite diln. is 0.71 = 0.01.

ARTHUR FLEISCHER

The status of the investigation of superconductivity. W. MEISSNER. *Metallwirtschaft* 10, 289-94, 310-5(1931), cf. *C. A.* 24, 5565.

E. II

Conductivity measurements in methyl ethyl ketone and acetone. P. WALLEN AND E. J. BIRN. *Z. physik. Chem. Abt. A*, 153, 1-51(1931), cf. *C. A.* 21, 1396, 24, 542.—The cond. of mono-, di-, tri- and tetra-alkylammonium picrates, iodides, bromides, chlorides, nitrates and perchlorates, of the picrates of Li, Na, K, Ag and Cd, of iodides of Na, K, Cd, and Hg, and of HgCl_2 were detd. at 25° in Me Et ketone. The alkyl groups include the Me, Et, n -Pr, n -Bu, isobutyl, isooctyl and n -octyl. The ketone had a cond. at 25° of 5.0×10^{-10} , d from 0° to 50° expressed by the equation $d = 0.82737(1 - 0.0012635t)$ and viscosity of 0.005220 at 0°, 0.003929 at 25° and 0.003115 at 50°. The cond. of NEt_4 picrate was also detd. at 0° and 50°. The conductivities of tetra- n -propyl, tetra- n -butyl and tetraisoamylammonium picrates and Li picrate were detd. at 25° in acetone. The limiting equiv. conductivities were calcd. by the square root law ($\Lambda_\infty = \Lambda + a\sqrt{c}$). Me Et ketone and acetone are differentiating solvents, as shown by the calcs. of the dissoc. consts. of the various salts. In the various solvents the order of decreasing dissoc. for the substituted NH_4 salts is the following: tetra-, tri-, di- and mono-. Kohlrausch's law of the independent migration of ions is substantiated. Stokes' law or the Walden rule of the correspondence of the temp. coeff. of mobility and of fluidity is confirmed by the measurements on NEt_4 picrate. Mobilities and solvation nos. were calcd. from the data. Substituted alkylammonium salts with 4 CH_3 groups give the following order of ion mobilities: $(\text{CH}_3)_4\text{N}^+ > (\text{C}_2\text{H}_5)_3\text{NH}^+ > \text{iso-C}_4\text{H}_9\text{NH}^+$, which is the order observed in alc. and in water. For the halogen and alkali ions, the mobility is inversely proportional to the size of the ion.

ARTHUR FLEISCHER

Acidity in non-aqueous solvents. Conductometric and electrometric titrations of acids and bases in benzene. VICTOR K. LA MEE AND HAROLD C. DOWNES. *J. Am. Chem. Soc.* 53, 888-90(1931).—Cond. titrations of acids and bases in C_6H_6 exhibit the anomalous behavior of salts (*C. A.* 23, 5088). Electrometric titrations of $\text{C}_6\text{H}_5\text{COOH}$ and $(\text{C}_2\text{H}_5)_3\text{NH}^+$ in C_6H_6 , with quinhydrone electrodes, resemble the titrations of aq. solns. of strong acids and bases.

CURTIS L. WILSON

New method for measuring the limiting potential of two phases. B. LANGEVEL.

Magyar Chem. Folyóirat 36, 89-91, 105-11 (1930).—The mutual charge of 2 outwardly isolated phases formed because of the change of potential in the double zone is used for quant measurement. The e/m of concn chain: normal electrode | soln | solid phase | soln | normal electrode, may be measured by this method. The method e_1 | dielectric | e_2 may be used directly for measuring the phase limit potential of quartz | aq. soln | even solid cryst. substances can be measured with slight modifications. The sensitivity of the method is ≈ 1.5 mv. A metal electrode was used instead of dielectric in control expts., results of which agreed with the data of the former method. Preliminary report of phase limit potential of quartz | electrolyte soln. Quartz glasses are examd on their II electrode function by measuring the e/m of an acid alkali-chain and an alkali acid chain. The reproducible difference totaled about 300 mv. Preliminary concn- e/m curves show maxima at pH 12. S. S. DE FÉNYI

Electrometric determinations of the dissociation of glycooll and simple peptides. PHILIP H. MITCHELL AND JESSE P. GREENSTEIN. *J. Gen. Physiol.* 14, 235-75 (1930).—The apparent acid and basic dissociation constants were detd potentiometrically by the methods of hydrolysis and titration for glycooll, glycylglycooll, alanylglycooll, valylglycooll, leucylglycooll, methylleucylglycooll, phenylalanylglycooll and glycylglycylglycooll. The dissociation constants were also detd in the presence of KCl and K_2SO_4 at equal ionic strength. The relative order of magnitude of the dissociation constants decreases with the increase in no. of C atoms between the NH_2 and $COOH$ groups. This is explained as follows on the basis of electronic structure with $AcOH$ and glycooll as examples. The highly electroneg. NH_2 group effects a deformation in the electronic nucleus causing a shift in position of the electron pairs toward the NH_2 group about the C and O atoms of the $COOH$ group. In the O—H linkage of the $COOH$ group, the electron pair is drawn closer to the O and farther from the H, decreasing the strength of the bond and increasing the degree of dissociation of the acid. The proton set free combines at once with the 2 unshared electrons of the N, forming the dipolar mol. or amphoteric ion. Therefore the farther away the NH_2 group is from the $COOH$ group, the smaller the deformation of the electronic nucleus and the slighter the dissociation. In general the dissociation constants were found to be functions of the II ion activity and of the ionic strength of the solns. Apparent contradictions to the Debye-Hückel theory are indicated and partially explained on the basis of specific ion effects. C. H. RICHARDSON

The electrode potential and the solvent: the solvation activity coefficient. FREDERICK K. V. KOCH. *Phil. Mag.* 11, 579-85 (1931), cf. *C. A.* 24, 6574.—The solvation activity coeff. may be defined by the equation $A_s = RT \ln a_s$, where A_s is the free energy involved in the transfer of an ion from pure solvent to a soln of specified concn. It is related to the usual activity coeff. by the expression $a = a_s a_o$, where a_o is the activity based on the transfer of the ion from vacuum to pure solvent. By considering the transfer of an ion from vacuum to soln through an electrode the equation for the electrode potential becomes $E = E_o - E_s + (RT/nF) \ln a_s$, where E_o is the solvation potential and E_s is the element potential. The abs. normal potential is thus the difference between two potentials, one dependent on the solvent and the ion, the other on the element electrode and the ion. The Nernst soln. tension is the distribution coeff. between pure solvent and the element electrode. Solvation potentials were calcd. for H , Li, Na, K, Rb, Cs, Tl, Ca, Cl, Br and I ions from the std. free energy of hydration. The solvation activity coeffs. were calcd. for ions in C_2H_5OH . The Born formula for the free energy of solvation holds only for the case where no sp. interaction of ions occurs. The Walden formula for solvation activities was not confirmed. ARTHUR FLEISCHER

Temperature coefficients of certain reference electrodes. PAUL PUTZEYS. *Bull. soc. chim. Belg.* 39, 454-7 (1930).—Expn. detns. are reported. For a 0.1 N HCl electrode, $E_t - E_{25} = 0.334 (t - 25) + 0.00279 (t - 25)^2$. For a 0.01 N HCl, 0.09 N KCl electrode, $E_t - E_{25} = 0.391 (t - 25) + 0.00335 (t - 25)^2$. A. L. HENNE

Experimental researches of potential differences between mercury, amalgam and dielectric. B. BICSEKOV. *Magyar Chem. Folyóirat* 37, 23-32 (1931).—Production of a zero electrode was tried by using different metal amalgams. Na, Zn, Cd, Sn, Pb and Ag were examd. without any result. S. S. DE FÉNYI

Glass-electrode measurements by means of a galvanometer with condenser attachment. MALCOLM DOLE. *J. Am. Chem. Soc.* 53, 620-2 (1931).—With a ballistic galvanometer as a null-point instrument measurements of glass electrode potentials were made to 0.1 millivolt and were shown to agree with those of a II electrode. M. D.

Quinhydrone electrode for mass experiments. G. HAYOS. *Magyar Chem. Folyóirat* 35, 168-70 (1929).—A practical electrode for mass detns. of pH is illustrated and described. S. S. DE FÉNYI

Determination of hydrogen-ion concentration by Haber-Klarnesiewicz' glass chain,

especially in soil examinations. G. HATOS AND GY. GOLL. *Magyar Chem. Folyóirat* 36, 33-40 (1930). —On examn. of 15 soils the method gave almost the same values for p_{H} as the detn. by the quinhydrone electrode. Agreement within about 0.1 p_{H} resulted from the use of a dil. aq. suspension. Thicker suspension gave agreement within 0.2-0.3 p_{H} . The quinhydrone electrode seemed to be reliable for p_{H} detns. S. S. DE FRÉJLY.

The measurement of the hydrogen-ion concentration in unbuffered solutions. II. Application of the hydrogen electrode. I. M. KOLTHOFF AND TOSI KAMEDA. *J. Am. Chem. Soc.* 53, 821-4 (1931). —It was found that an electrode made by crating a Pt wire with a thin bright layer of Pt yields satisfactory results in the measurement of p_{H} in unbuffered or slightly buffered solns. The p_{H} values of $(\text{NH}_4)_2\text{SO}_4$, ZnSO_4 , and KCl solns were measured. III. The colorimetric method. *Ibid.* 825-32. —The measurement of the p_{H} of pure water, of neutral salts in pure water, and of very dil. NaOH solns by the colorimetric method is described. The accuracy of the isohydric indicator method is 0.05 p_{H} in pure water and 0.1 p_{H} in the NaOH solns. CHANNING WILSON.

Determination of hydrogen-ion concentration with a photoelectric colorimeter. RALPH H. MILLER AND HERMAN M. PARTRIDGE. *Ind. Eng. Chem., Anal. Ed.* 3, 169-71 (1931). —The operation of the colorimeter depends upon the variation in photoelectric current with the transmittancy of the indicator soln. with changing p_{H} value. The feeble photoelectric current is supplied to a 3-electrode vacuum tube of suitable mutual conductance and p_{H} values are read in terms of the plate current of the triode. A filter is used between the light source and the soln., being such that it transmits light of a wave length corresponding to one absorption max. of the soln. The plate currents obtained with solns. contg. the same amt. of indicator but buffered to various p_{H} values are plotted against p_{H} , giving the calibration curve for the indicator. If all factors are const., the plate current of the unknown soln. (contg. the same amt. of indicator) gives the p_{H} . An alternative method is to adjust the intensity of the lamp, vacuum tube filament, etc., to the same value for each expt. and to multiply the observed plate current by the factor representing the ratio between p_{H} and the current for the indicator. Const. results were obtained over long periods of operation. Diagrams of the app. and calibration curves for various indicators are given, and the characteristics required in suitable vacuum tubes discussed in detail. EDUARDO B. SANTIAGO.

The chemical reactivity of the fused bases. I. The action of the alkali amides upon electropositive metals. W. CONARD FERNELIUS AND F. W. BERGSTROM. *J. Phys. Chem.* 35, 740-55 (1931), cf. *C. A.* 22, 1519. —An app. is described for the study of reactions of pure fused alkali amides which avoids contamination of the melt and permits viewing the reaction, controlling the temp., altering the fusion stms., quant. collection of gases and extn. of the melt with liquid NH_3 . Mg reacts with NaNH_2 or KNH_2 in an NH_3 atm. to liberate the alkali metal. $\text{Mg} + 2\text{KNH}_2 \rightleftharpoons \text{Mg}(\text{NH}_2)_2 + 2\text{K}$, the equl. point being far to the right. Subsequent reactions are $\text{Mg}(\text{NH}_2)_2 + 2\text{KNH}_2 \rightarrow \text{Mg}(\text{NHK})_2 \cdot 2\text{NH}_3$; $2\text{NH}_3 + 2\text{K} \rightarrow 2\text{KNH}_2 + \text{H}_2$. Ca and Al also liberate K . Ca forms $\text{CaNK} \cdot 2\text{NH}_3$ and Zn forms $\text{Zn}(\text{NHK})_2 \cdot 2\text{NH}_3$, while Al forms $\text{AlN} \cdot \text{NH}_3$ + adsorbed KNH_2 . Be forms $\text{BeNNa} \cdot 2\text{NH}_3$ with NaNH_2 . Ce , Th and Mn are slightly attacked by fused KNH_2 , while Hg , Cu , Cd , Ti , Ta , Zr , Ta , Cr , Ni , Pt and Ir are not noticeably attacked. Na gives with fused NaOH a blue color, Mg gives a similar color suggesting the reactions $\text{Mg} + 2\text{NaOH} \rightleftharpoons \text{Mg}(\text{OH})_2 + 2\text{Na}$, $\text{Mg}(\text{OH})_2 + 2\text{NaOH} \rightarrow \text{Mg}(\text{OH})_2 + 2\text{H}_2\text{O}$, $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ (cf. *C. A.* 4, 871, 7, 2167). Ca and Na appear to give some hydride upon reaction with NaOH . W. C. FERNELIUS.

Iron oxides and their reducibility with hydrogen and carbon. F. YA. RODE. *J. Russ. Phys.-Chem. Soc.* 62, 1453-66 (1930). —The heating curve of Fe_2O_3 has a conversion point at 745° , quite likely corresponding to a definite modification of Fe_2O_3 . Beginning with 1280° some irregularities are observed that probably have a connection with the beginning of loss of O . The temp. curve of reduction of Fe_2O_3 with H_2 has 2 kinks probably corresponding to 2 steps in reduction, first to Fe_3O_4 and then to a lower oxide. There is no definite proof of formation of FeO as an intermediate product. Fe_2O_3 is more easily reduced with H_2 than Fe_3O_4 . In reducing hydrated iron oxides with H_2 , the point where reduction begins could not be ascertained, because it lies in the region of decompn. of corresponding hydrates. If the products of reduction are cooled in H_2 , they become pyrophoric when the heating temp. does not exceed $600-650^\circ$. A higher heating temp. makes them non pyrophoric. Reduction of Fe_2O_3 and hematite ore with C gives on the temp. curve 3 halts, probably corresponding to 3 consecutive stages of reduction to the metal. Here again Fe_2O_3 is more easily reduced than Fe_3O_4 . S. L. MADORSKY.

The reaction mechanism of carbon combustion at low pressures. V. SHILOVICH

Z. Elektrochem. 36, 806-7(1930) —The combustion of metallized C filaments was carried out in pure O as well as in additional gases. The reaction velocity curves for the formation of the combustion products in pure O are expressed in micromol. per mm. O pressure cm^2 surface and second. The results may be summarized as follows: Kinetic measurements on the C combustion at low O pressures, in general indicate a reaction of the first order. The combustion in streaming O between 800° and 1400° forms a gas mixt of const compn ($\text{CO}_2 + 2\text{CO}$). Between 1400° and 1500° the C surface suffers a change whereby, according to the observation of Langmuir, CO is vaporized. Here the velocity curves of both combustion products of C indicate a break. Between 800° and 1200° CO formation is promoted through autocatalysis. The presence of CO_2 retards the combustion of CO and C. The velocity equation of CO_2 in the presence of CO for the temp. between 1500° and 1600° gives an energy value of approx -17 kg.-cal. At 1600° a minute addn of water vapor accelerates the formation of both combustion products catalytically. It appears that the adsorption heat of O on the C of approx. 49 kg. cal. observed by Marshall and Bramston-Cook, corresponds to the formation of an atomic adsorption layer of O already proposed by Langmuir. W. VANSELOW

Problems and results of the newer investigations of free radicals. KARL ZIEGLER *Z. anorg. Chem.* 43, 915-9(1930) —This is a concise discussion of the causes that det. whether or not a particular substance will be strongly or weakly dissociated. The principal points considered are the energy of activation of C_2H_2 and the radical and steric factors. A. L. HENNE

Kordes' law. K. M. STAKHORSKII *J. Russ. Phys.-Chem. Soc.* 62, 1493-7(1930) —Aside from normal cases where both component compds of a system are unassociated Kordes' law can be used to det. approx. from a study of the m-p diagram the mol condition of pairs of which one or both compds are assocd. In such cases the assocn const. X_a is detd. by placing $X_a M_a$ for M_a in Kordes' equation, when A is the assocd compd. The equation then becomes $X_a = [M_a a(T_a - T_e) T_b / M_b b(T_b - T_e) T_a]$, where a and b are consts of the 2 components A and B , M_a and M_b are their mol wts., and T_a , T_b and T_e are the m. p.s. of A , B and the eutectic mixt. The fact that H_2O , alics. and org. acids are typical assocd compds is thus once more confirmed by the application of Kordes' law. S. L. MADORSKY

Relation between energy of activation and constant S of Arrhenius' equation. E. N. GAPON *Ukrainskii Khim. Zhur.* 5, Sci. Pt., 169-80(1930) —By the example of 4 reaction groups, it is shown that in a given group the reactions are characterized by the so-called inversion temp. T_i , at which the reactions proceed at the same rate. From this fact results the relationship between the energy of activation E and the const. S of Arrhenius' equation $\log S = \log K_i + (E/RT_i)$, where K_i is the rate of reaction const. at T_i . Thus K_i and T_i are connected by the equation $\log K_i = \log \psi + (\varphi/RT_i)$, where ψ and φ are consts., characteristic of all reactions. An equation is given describing the kinetics of 28 monomol. reactions. S. L. MADORSKY

The problem of dependence of yield on temperature. E. N. GAPON *Ukrainskii Khim. Zhur.* 5, Sci. Pt., 165-7(in German 168)(1930) —Starting from equations of uni- and bimol. reactions and the equation of Arrhenius, 2 equations are obtained showing the relation between temp. and reaction yield $\log\{\log A/(A-x)\} = a_1 - (b_1/T)$ for monomol. reactions, and $\log\{X/(A-x)\} = a_2 - (b_2/T)$ for bimol. reactions, where x = reaction yield at a given time, A = initial concn., T = abs. temp. and a_1 , b_1 , a_2 , b_2 are consts. These equations make it possible to det. the order of reaction and activation energy from yield data at various temps. S. L. MADORSKY

Equilibrium constants for the decomposition of ammonium bicarbonate. Wm. K. HUTCHINSON *J. Chem. Soc.* 1931, 410-2 —The equil. consts. for the decompn. of NH_4HCO_3 are calcd. from detns. of the partial pressures of NH_3 , CO_2 and H_2O in equil. with satd. NH_4HCO_3 solns. The mean values of K and K_1 (detd. by assuming the partial pressure of H_2O vapor to be const. at const. temp.) at 10° and 20° are:

$$K = p_{\text{CO}_2} \times p_{\text{NH}_3} \times p_{\text{H}_2\text{O}} \quad K_1 = p_{\text{CO}_2} \times p_{\text{NH}_3}$$

At 10°	253	29.6
At 20°	3000	180

The general equation for K is $\log_2 K = 33.88 - 40,600/2.303RT$. J. BALOZIAN
Equilibrium in the iron-oxygen-hydrogen system at temperatures above 1000°. WALTER E. JOMINY AND DONALD W. MURPHY. *Ind. Eng. Chem.* 23, 384-7(1931) —A method was developed to det. the equil. point by regulating the ratio of H_2 to water vapor in gas passing over the Fe sample at a given temp. The equil. was detd. from 1097° to 1427°. From these data the m. p. of FeO is given as 1357° and the heat of fusion as +29,000 = 5000 cal. per g. mol. B. E. TIFFANY

Dissociation of sulfuryl chloride KIMIO ARII *Bull Inst Phys-Chem Research (Tokyo)* 10, 256-65(1931), (Abstracts 32-3 [in English] published with *Sci Papers Inst Phys Chem Research (Tokyo)* 15, Nos 289-91) — K , was measured for satd and unsatd vapors of carefully purified SO_2Cl_2 at 30°, 40°, 60°. The mean values agree with the calcd ones. ALFRED BURGER

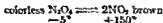
Dissociation of silver carbonate MITSUKI WATANABE *Bull Inst Phys-Chem Research (Tokyo)* 10, 241-50(1931), (Abstracts 29-30 [in English] published with *Sci Papers Inst Phys Chem Research (Tokyo)* 15, Nos 289-91) —The pressure of dissociation of Ag_2CO_3 was measured by a static method between 120° and 201°. The values obtained agree with the calcd ones. The following thermodynamic quantities were calcd: ΔH_{298}° for the formation of Ag_2CO_3 , -119 962 cal, ΔF_{298}° -103 614 cal, $S_{298}^\circ(\text{Ag}_2\text{CO}_3)$ at 25° 42.5 cal/degree. ALFRED BURGER

Investigation of equilibrium between vapor and solution in the system: acetic acid-acetic anhydride. S. I. LUGANOV *J Russ Phys-Chem Soc* 62, 1509-21 (1930) cf *C A* 24, 4451 —The degree of assocn of HOAc in soln depends on the nature of the other component and on temp. A study was made of the vapor pressure of HOAc in the system $\text{HOAc} - \text{Ac}_2\text{O}$ at 60° and 80°, by M. Vrevskii's exptl method. From these data the degree of assocn of HOAc in the soln was calcd by means of Vrevskii's and Gibbs' equations. S. L. MADORSKY

The study of the complex silver hyposulfites of sodium E. CASATKÉ AND RAULNY *Compt rend* 192, 423-5(1931) —The reaction of AgCl with $\text{Na}_2\text{S}_2\text{O}_3$ solns was studied by applying the mass law. The $\text{Na}_2\text{S}_2\text{O}_3$ is added from a buret to a known quantity of AgCl until completely dissolved. The ionic reaction is $\text{AgCl} + \text{S}_2\text{O}_3^{2-} \rightleftharpoons \text{AgS}_2\text{O}_3^- + \text{Cl}^-$. Since the concns of these ions are known, K can be calcd. An excess of NaCl has little effect on K , but AgNO_3 in excess forms other complex compds and could not be studied in this way. From the increase in K with temp., the heat of formation of the complex ion was detd as 7.7 large cal. The cond of the mixt. after the soln of the AgCl is the same as that of the original soln since the number of ionic charges does not change. ARMY LEVESONTE

The equilibrium between chlorine, bromine and bromine chloride (BrCl). W. JOST *Z physik Chem, Abt A* 153, 143-52(1931) —In the extinction curve for BrCl absorption is 0 at 575 $\text{m}\mu$ so that measurements beyond this range are not affected by BrCl absorption. Measurements were made of the correction necessary because of absorption of Br by foreign gases. The expts at 18°, 88° and 230° yield the equation $\log K = (1320/4.57T) - 0.18$. The heat of formation of BrCl calcd is 0.75 = 0.5 kg.-cal. ARTHUR FLEISCHER

Thermal dissociation of nitrogen peroxide. O. KLAUSVIC *Chem Listy* 24, 473-4 (1930) —A glass tube 12-15 cm long and 1 cm in diam is closed at one end and partly filled with powd $\text{Pb}(\text{NO}_3)_2$ dried at 100°. The tube is sealed at the other end, and heated carefully and in intervals until the reaction $2\text{Pb}(\text{NO}_3)_2 \rightleftharpoons 2\text{PbO} + 4\text{NO}_2 + \text{O}_2$ is completed. The tube is filled with brown NO_2 and carefully cooled. By plunging the tube into cold H_2O , the contents become almost colorless as a result of the formation of N_2O . Warming the tube yields NO_2 again. The equation for the dissociation is



FRANK MARESH

The interpretation of the thermal decomposition of nitrous oxide. H. C. RAMSPERGER AND G. WADSWORTH *Proc Natl Acad Sci* 17, 103-5(1931) —The thermal decompn of N_2O has recently been found unimol (cf *C A* 24, 4980, 25, 893). According to the theories of Kassel and Rice and Ramsperger, only 1 classical oscillator is required to fit the data at 665°. The mol diam. is 3.05×10^{-8} cm., the crit energy, 53,000 cal per mole and $k_\infty = 1.78 \times 10^{-12}$. The const. A in the equation for the rate const is 4.0×10^6 . A table is given for the 11 known unimol reactions giving the pressure and temp ranges that have been studied, the consts A and Q in the rate const equation and the no of classical squared terms required by the theory. G. M. M.

The thermal decomposition of nitroxyl chloride. A homogeneous gas reaction of the first order. HANS-JOACHIM SCHUMACHER AND GERRARD SPRENGER *Z physik Chem, Abt B*, 12, 115-31(1931) —A description and sketch of the app for measuring thermal decompn of NO_2Cl is given together with the procedure for collecting the accompanying data. It is shown that the reaction is monomol and is still independent of pressure at 10 atm. Its heat of activation is calcd to be 20.5 kg.-cal. All contaminating gases increase the rate of decompn about equally while NO_2 does so about 20% faster. Stoichiometric amts of NO_2 and Cl_2 act like unchanged NO_2Cl . The exptl data are discussed. R. H. LAMBERT

The thermal decomposition of chlorine monoxide. J. J. BEAVER AND G. STIEGER. *Z. physik. Chem., Abt. B*, 12, 93-108(1931).—The thermal decomposition of Cl_2O follows a complicated chain reaction similar to that of ClO_2 , rather than the simple bimolecular reaction believed to occur. Between 100° and 130° the rate is proportional to monoxide concentration and at 140° to the exponent 0.75 of concentration. A consistent explosion was observed at the end of the reaction. A mechanism of the reaction agreeing with the experimental results concludes the article. R. H. LAMBERT

The influence of foreign gases in gas decomposition reactions. N. NAGASAKO. *Z. physik. Chem., Abt. B*, 11, 420-4(1931).—The theory of monomolecular reactions is extended to include the presence of foreign gases. New observations on the decomposition of N_2O in the presence of H_2 are reported. J. B. AUSTIN

Kinetics of the combination of hydrogen and oxygen: Influence of iodine. W. L. GARSTANG AND C. N. HINSHLEWOOD. *Proc. Roy. Soc. (London)* A130, 640-54(1931).—The upper limit of the low pressure explosion region in the $\text{H}_2 + \text{O}_2$ reaction is lowered slightly by pretreatment with I_2 , it is lowered also by water vapor in amounts up to 2.7% of total pressure and above this percentage explosion is entirely inhibited. The surface reaction in SiO_2 vessels is accelerated by I_2 . The gas reaction at high temperatures and pressures is retarded by I_2 and the explosion in the critical low pressure region does not take place with I_2 in concentration greater than a minute amount. Amounts of I_2 less than required to stop explosion entirely lower the explosion pressure. S. LEVNER

The kinetics of the reduction of cuprous oxide. I. The reduction at low pressures—with an appendix giving the results of further experiments on the oxidation of copper at low pressures. FREDERICK J. WILKINS. *J. Chem. Soc.* 1931, 330-5.—The reduction of Cu_2O by H_2 is studied at pressures lower than 0.2 mm. under such conditions that the $\text{Cu-Cu}_2\text{O}$ interface is effectively constant and the H_2O formed is instantly removed from the reaction vessel. The rate of reduction is proportional to the pressure of H_2 . The heat of activation in the reduction process is 18,000 cal. as judged by measurements between 144° and 250° , it is to be identified as the heat of evaporation of the adsorbed H_2 into Cu_2O . Earlier experiments on the rate of oxidation of Cu are extended to low pressures, the oxidation rate is proportional to the pressure of O_2 and involves an energy of activation of 9000 cal. P. H. EMMETT

Chemical kinetics of ion reactions. I. Principles of the theory of Brønsted. Á. KISS. *Magyar Chem. Folyóirat* 35, 139-44(1929).—Short summary of Brønsted's theory and its applications. II. Mechanism of reaction between ferric and iodide ions. *Ibid.* 36, 49-59(1930).—The reaction proved to be bimolecular for I^- ions and monomolecular for Fe^{+++} ions. Hydrolyzed ferric ions do not react with I^- ions. The primary kinetic salt effect agrees with Brønsted's neutral salt law in dilute solutions. Alkali metals show a striking specific ion effect. Slowing influence of ferrous and sulfate ions must be considered as a secondary kinetic salt effect. III. Neutral salt effect and catalysis in case of ion reactions. *Ibid.* 37, 17-23(1931).—Connections between the theory of Brønsted and catalysis, kinetic salt effect and medium effect are discussed. S. S. DE FINALLY

Kinetics of chlorine bleaching. JOSEPH J. WEISS. *Z. anorg. Chem.* 44, 102-4(1931); cf. *C. A.* 25, 2044.—Polemical with H. Kauffmann (*C. A.* 19, 561; 24, 5589) H. KAUFFMANN. *Ibid.* 104-5(1931), cf. *C. A.* 24, 5589, 25, 597.—A reply. E. M. SYMMES

Reaction velocity in the system $\text{Ag}_2\text{CO}_3 \rightleftharpoons \text{Ag}_2\text{O} + \text{CO}_2$. WILFRED D. SPENCER AND BRYAN TOPLEY. *Trans. Faraday Soc.* 27, 94-102(1931); cf. *C. A.* 24, 2366.—Previous work was extended to include the recombination of Ag_2O and CO_2 , and the decomposition of the Ag_2CO_3 at various pressures. When the pressure of the CO_2 is several times as great as the equilibrium pressure, recombination takes place rapidly at first and then the rate decreases suddenly. This initial rapid reaction becomes slower as the pressure is decreased. The decomposition of the synthetic Ag_2CO_3 in a vacuum is more sensitive to temperature than that of the crystalline H_2O vapor promotes recombination by facilitating the penetration of CO_2 through the Ag_2CO_3 first formed. CURTIS L. WILSON

Salt and medium effects on the temperature coefficient of the velocity of decomposition of diacetone alcohol. GEO. M. MURPHY. *J. Am. Chem. Soc.* 53, 977-81(1931).—The velocity of decomposition of diacetone alcohol was measured by a dilatometric method similar to that of Åkerlöf (*C. A.* 21, 688). Different concentrations of aqueous NaOH solutions were used, also aqueous NaOH-NaCl solutions and aqueous MeOH mixtures containing NaOH. The temperature coefficient of reaction velocity is independent of the concentration of NaOH and NaCl in the NaOH-NaCl solutions. In the MeOH-NaOH solutions the temperature coefficient increases with increasing MeOH content. The energy of activation was computed as 18,000 cal. E. J. R.

Chemical kinetics in mixtures of solvents. VIII. G. E. MURKIN AND R. B. GINZBURG. *Ukrainskii Khim. Zhur.* 5, No. 2, Sci. Pt. 147-58(1930), cf. *C. A.* 24, 2042.—

The velocities of reactions were studied between $\text{C}_6\text{H}_5\text{N}$ (I) and allyl bromide (II) and between PhNMe_2 and II in mixts of Me_2CO and CHCl_3 at various temps. The observed consts of reaction velocities of the mixts are smaller than the values calcd. according to the additive rules. An increase in temp enlarges the percentage deviation of the observed consts from the additivity. Under the same conditions this deviation is greater between PhNMe_2 and II than between I and II. IX. G. I. MUKHIN AND R. I. BARANOVA *Ibid* 159-62.—The velocity of reaction between I and II was studied in equimol concns of 0.5 g mol per l in AcOEt , $\text{MeCH}_2\text{CO}_2\text{Et}$, $\text{MeC}_2\text{H}_4\text{CO}_2\text{Et}$ at 35° and 45° , in BzOMe , BzOEt and BzOPr , and $\text{C}_6\text{H}_5(\text{OH})\text{CO}_2\text{Me}$ and $(\text{CO}_2\text{Et})_2$. The most suitable solvents for this reaction are the esters of BzOH . The greater the mol wt of the esters of fatty acids and BzOH , the smaller is the velocity const of the reaction. The introduction of OH (BzOH to $\text{C}_6\text{H}_5(\text{OH})\text{CO}_2\text{Et}$) increases the velocity const.

C. H. BLANC

The rate of solution of metals in acids. Contribution to the theory of local voltaic elements. W. GLEISLER AND B. BLUMENTHAL. *Z. physik. Chem.*, A, 152, 197-234 (1930).—The general course of the soln curves of Fe, Ni, and Cu in 0.1 N HCl and the detailed rate of soln curves of 2 different samples of Zn were detd. Exptl conditions were found under which reproducible values could be obtained. The rate of soln of Zn in 0.2 N HCl compared to 0.1 N HCl is not even approx proportional to the conductivities of the 2 acids, contrary to the results of Fricson, Auren and Palmaer (*Z. physik. Chem.* 39, 1 (1902), 45, 152 (1903), 56, 689 (1906)). In general, the rate of soln of metals in acids is proportional to the product of the cond. of the acid soln and the difference between the anode and cathode potentials of the local voltaic element, Zn is the anode in the present case, whereas various impurities present in the metal constitute the cathode of the so-called local element. The observed acceleration of the rate of soln with time in a particular acid is interpreted as being due to a decrease in the H₂ overvoltage of the cathode. When this observed variation of H₂ overvoltage is taken into account, the theory of local elements satisfactorily explains the rate of soln of Zn in HCl.

P. H. FRETWELL

The inhibition of chemical reactions. IV. The site of antioxygenic action. The oxidation of sodium sulfoxide and benzaldehyde. KENYLTN C. BAILEY AND VIOLET H. FFRYCH. *J. Chem. Soc.* 1931, 420-3, cf. *C. A.* 24, 2947.—The inhibition of the oxidation of aq. Na_2SO_3 solns by isopropyl and *sec* butyl alc. in aq. soln is studied, the method of McBain and co-workers (*C. A.* 21, 3510, 24, 1560) being used. For concns of the alc. in the aq. soln greater than 2 g/l, Γ (increased concn of the alc. in the surface layer in g/sq cm) has a mean value of about 8.6×10^{-4} g/sq cm for isopropyl and 10.3×10^{-4} g/sq cm for *sec* butyl alc. and varies only slightly with change in concn. The results indicate an abrupt change in Γ between 2 and 1 g/l, the unimol layer being formed at a concn. of about 1.5 g/l. It is shown that the inhibitor is preferentially adsorbed at the interface in aq. solns of those alics that inhibit the oxidation of the Na_2SO_3 solns. The graph of the oxidation velocity of aq. Na_2SO_3 soln against the inhibitor (alc.) concn. does not show a discontinuity at this concn. of alc. That the autooxidation of BzH is not identical with that of aq. Na_2SO_3 solns is suggested by the velocity inhibitor curves.

J. BALOZIAN

The molecular constitution of the α solid solutions of tin in copper and of the corresponding liquid solutions examined thermodynamically. F. H. JEFFERY. *Trans. Faraday Soc.* 27, 136-7 (1931), cf. *C. A.* 24, 2660, 25, 632.—From the equation used in previous work, it is concluded that the α solid soln, like the liquid soln, consists of Cu₃Sn dissolved in monat Cu.

CURTIS L. WILSON

The molecular constitution of the β solid solutions of tin in copper examined thermodynamically. F. H. JEFFERY. *Trans. Faraday Soc.* 27, 137-9 (1931), cf. preceding abstract.—Evidence based upon the previous thermodynamic equations indicates that the β solid solns are monat Sn dissolved in monat Cu.

CURTIS L. WILSON

A quantitative relation between the slope dP/dT of the curves representing equilibrium at an invariant point. (MILLER) A. E. KORVETZ (MILLER) N. H. J. M. VOOGE AND P. F. C. SCHEFFER. *Rec. trav. chim.* 50, 252-5 (1931).—Equations are derived for calcg. dP/dT or the compn. of the unknown phase at an invariant point (cf. following abstr.).

G. M. MURPHY

The composition of the hydrate of hydrogen sulfide. (MILLER) A. E. KORVETZ AND P. F. C. SCHEFFER. *Rec. trav. chim.* 50, 256-60 (1931), cf. *C. A.* 13, 3049 and preceding abstr.—Data of the slope of 3 P - T curves and the compn. of the phases at the quadruple point for the system H_2S - H_2O gives $\text{H}_2\text{S} \cdot 6\text{H}_2\text{O}$ as the compn. of the hydrate.

G. M. MURPHY

Thermodynamic relations in multi-component systems. ROY W. GORANSON

Carnegie Inst Washington Pub 408, 310 pp (1930) — A complete outline of the subject with accompanying tables

JOSEPH S. HIRSHURN

The equilibrium diagram of the chromium-carbon system. KAZUO HATSUTA *Kinzoku No Kenkyu* (J. for Study of Metals) 8, 81-8(1931) — The Cr-C system was studied by means of thermal, microscopic and x-ray analyses and the equilibrium diagram is presented. There are 4 carbide phases, ϵ , η , δ and α corresponding to compounds Cr_7C_3 , Cr_3C_2 , Cr_7C_3 and CrC (δ), resp. The crystal form of Cr_7C_3 is cubic, that of Cr_3C_2 trigonal and Cr_7C_3 orthorhombic, but that of CrC was not determined. α - and ϵ -phases form a eutectic at 1485° , 3.7% C. The ϵ - and η phases are peritectically formed according to the following scheme, resp. $\eta + \text{melt} \rightarrow \epsilon$, $\delta + \text{melt} \rightarrow \eta$. The peritectic temp. of the former is 1530° , and that of the latter is above 1600° . The δ -phase probably has the max. m.p. on the liquidus curve and forms a eutectic with α -phase, a transformation takes place at 1505° in this phase.

M. KURODA

Fractional crystallization in the Blanc process with hydrochloric and nitric acids. GIOVANNI MALQUORI *Atti III congresso nat. chim. pura applicata* 1930, 420-31, cf. *C. A.* 23, 2876 — Discussion of various ternary systems of the 5-component systems $\text{AlCl}_3\text{-FeCl}_3\text{-KCl}$, HCl , H_2O and $\text{Al}(\text{NO}_3)_3\text{-Fe}(\text{NO}_3)_3\text{-KNO}_3\text{-HNO}_3\text{-H}_2\text{O}$ which are of interest in connection with the extn. of K and Al salts from leucite by the Blanc process. Consideration of the phase diagrams shows how large yields and extreme purity of the products may be obtained.

G. M. MURPHY

Melting point diagram in the system zirconium dioxide beryllium oxide. O. RUFF, I. FREAT and H. V. WARTENBERG *Z. anorg. allgem. Chem.* 196, 335-6(1931), cf. *C. A.* 24, 3330 — Earlier expts. are repeated and the conclusion is reached that no compound is formed in the system $\text{ZrO}_2\text{-BeO}$, but that there is a eutectic line at about 2200° . In the system $\text{ZrO}_2\text{-CaO}$, the compound ZrCaO_2 is demonstrated, with m.p. 2350° .

H. S.

Autocatalysis in oxidation. III. The mechanism of the action of positive catalysts in the autoxidation of abietic acid. G. DUPONT, J. LEVY and J. ALLARD *Bull. soc. chim.* 47, 912-52(1930), *Bull. inst. chim.* 1931, 20-30, cf. *C. A.* 24, 855, 2915, 25, 1148 — The autoxidation of abietic acid is a normal autocatalysis, normally, it is greatly accelerated by Co abietate, acting as a pos. catalyst. However, an extremely small quantity of catalyst (about 0.001 of Co salt) has a marked antioxygen effect. O is absorbed rapidly, at an increasing rate, until one atom has been absorbed by the abietic acid, the oxidation continues then much more slowly, until a second atom of O has been taken up. From there on, the absorption of O becomes an extremely slow process. The results are explained by the Moureu-Dufraisse theory. It is claimed that the effective catalyst is probably a complex oxidized acid Co abietate, the supporting evidence is furnished by a comparison of the absorption spectra of slightly acid Co abietate, Co abietate in abietic acid and oxidized abietic acid contg. Co abietate.

A. L. HENNE

Cysteine and glutathione as anticatalysts in oxidations with molecular oxygen. A. SCHÖNBERG *Ber.* 64B, 546-61(1931) — The rate of oxidation of leuco methylene blue by mol. O_2 in the presence of $\text{Cu}(\text{OAc})_2$ was found to be due to 2 reactions: a primary involving the transfer of O_2 by the $\text{Cu}(\text{OAc})_2$, and a secondary in which the leuco dye is oxidized by the H_2O_2 in the soln. The amt. of O_2 taken up by the leuco dye depends upon the ratio of the velocities of these 2 reactions. The addn. of cysteine (I), cysteine hydrochloride (II) or glutathione (III) retards the oxidation. Curves are given showing the effects of the addn. of various concns. of I, II or III on the velocity of the reaction. In the presence of I, II or III the reaction velocity increases with time, while in their absence it decreases. This increase is thought to be due to the oxidation of I, II or III by H_2O_2 to the disulfides, which are considered inactive, for cysteine has only a very weak anticatalytic action. Increasing the concn. of I, II or III retards the oxidation, which is probably due to the formation of complex Cu compounds. In high concns. of I, II or III the Cu complex is only slightly dissoed, while in low concns. there is greater dissoen. and hence a higher concn. of the catalytically active Cu ions. HCN , which also forms Cu complexes, retards the oxidation of leuco methylene blue with mol. O_2 in the presence of $\text{Cu}(\text{OAc})_2$, but has no effect on the autoxidation of the leuco dye. Since the addn. of cysteine causes little retardation, its Cu complex may be more highly dissoed. It is also possible that the SH group may have a sp. effect.

RUTH BERGGREN

A simple method of measuring the specific heat of a solid body at ordinary temperature. Application to beryllium. P. VERNOTTE and A. JEUPROY. *Compt. rend.* 192, 612-4(1931) — A method is very briefly described for detg. sp. heat of solids without the temp. uncertainties encountered in the usual procedure. The solid is allowed to come to a const. temp. in a thick, tightly closed metal box and is then dropped into cold water in the calorimeter and the reading is recorded. This method can be used for higher

temp. in which case an oven is used to regulate the temp. Be was found to have a ΔH of 0.39° at 13° J. R. ADAMS

Vapor pressure and heat of dilution. VIII. Heat of dilution of cane sugar in aqueous solution and of urea and calcium chloride in alcoholic solution. R. B. VALLENDER and E. P. PERMAN *Trans. Faraday Soc.* 27, 124-35 (1931), cf. C. A. 25, 2350. The heats of diln. of cane sugar were measured at 0°, 20°, 40°, 60° and 80°, over a range of concn. from near satn. down to about 25%. Small values, never greater than 3 cal. per g. of H_2O , were obtained for all heats. They were all neg. Urea in alc. gave very small heats of diln., all pos. $CaCl_2$ in alc. had neg. heats of diln. The heats of diln. calcd. from the vapor pressures agreed with the exptl. values for urea and $CaCl_2$, but not for cane sugar. The osmotic pressures were calcd. from the heats of diln. of cane sugar, and agreed with the exptl. values. CRISTIS L. WILSON

The heat of formation of gaseous mercury, cadmium and zinc halides. H. SPONER *Z. physik. Chem., Abt. B*, 11, 425-32 (1931).—It is assumed that the bands of the diatomic Hg, Cd and Zn halides reported by Wieland (cf. C. A. 23, 4588) represent transitions from the basic level to an excited state. The vibration quanta were extrapolated on this basis and give the following heats of dissociation (in kg.-cal.) $HgCl$ 32, $HgBr$ 25, $CdBr$ 60, CdI 32 and ZnI 46. By combining these results with existing thermochem. data the following are obtained: $HgCl + Cl = HgCl_2 + 72$ kg.-cal.; $HgBr + Br = HgBr_2 + 62$ kg.-cal.; $CdI + I = CdI_2 + 48$ kg.-cal. J. B. ACSTON

The heat of formation of the nitrides. I. Manganese nitride and chromium nitride. B. NEUMANN, C. KRÖGER and H. HÄBLER *Z. anorg. allgem. Chem.* 196, 65-75 (1931).—A new method is described for detg. the heat of formation of nitrides directly by heating the metal in a small elec. furnace, contained in a bomb calorimeter, in N_2 under pressure. One mm., 45 sec. sufficed to heat the metal to 900° and nitride it. The Mn nitride formed at 500° to 1000°, under 10 to 25 atm. N_2 , is closely Mn_3N_4 ; and the heat of formation is, $5Mn + N_2 = 3Mn_3N_4 + 57,180$ (± 400) cal. By use of the less accurate method of burning nitrided Mn in O_2 in a bomb calorimeter the value, 62,400 cal., was detd. When Cr was nitrided at 600°, 25 atm. N_2 , for 1 min., 15 sec., only 0.65% N was absorbed, but at 900°, in 1 min., 45 sec., 15 to 24 g. N per 100 g. Cr reacted. The product was largely Cr_3N . The heat of reaction of 1 at. of N with Cr was found to be 29,500 \pm 500 cal. The exptl. detd. heats of formation agree poorly with values calcd. from dissociation pressure data, indicating that the dissociation pressures in these systems, metal N_2 , are not sufficiently accurate to permit the estn. of the influence of mixed-crystal formation on the quant. calen. of the equil. by use of thermodynamic equations. The electrolytic prepn. of pyrophoric, metallic Mn and Cr, with a Hg cathode, is described. R. H. LOWARD

The reversible oxidizability of organic substances: thermochemistry of the oxidation of rubrene. CHARLES DUTRAISSE and LEON ENDELLIN *Compt. rend.* 191, 1321-3 (1930).—The reversible oxidation of rubrene ($C_{40}H_{36}$) appears similar to that of hemoglobin. The photochem. oxidation of rubrene is exothermic, indicating that light acts as a catalyst and that other catalytic agents must exist. The heats of formation of rubrene and its oxides in Cals. are: Rubrene (R)—131, (oxyrubrene (dissociable), (R-O $_2$))—108.4, isooxyrubrene (RO $_2$)—50.4, metrubrene (RO)—92.4. The liberation of 23 Cals. in forming the dissociable oxide is much less than the heat (about 50 Cals.) evolved by introducing O into ethylene and the heat of formation of the other rubrene oxides. H. W. WALKER

van't Hoff's stability rule. B. BRICE *Rec. trav. chim.* 50, 230-7 (1931).—Available exptl. data support the rule that from 2 modifications of a system, the one with higher sp. heat becomes stable at higher temp. Ordinarily low temps. suffice for application of this rule but in case of intersecting (gas-solid) sp.-heat curves lower temps. must be considered. The rule also fails for cases like ortho- H_2 , where the modification appears only due to the lowering of the thermodynamic potential of the other form. G. M. MURPHY

The transformation of heat into work and "chemical" energy into heat as viewed in the light of the modern "volume energetics." RICHARD V. DALLWITZ-WAGNER *Wärme* 52, 549-50, *Chem. Zentr.* 1929, II, 1385.—If the abs. temp. is replaced by $(1/\delta)$, δ representing the cubic coeff. of expansion at the resp. temp., then the thermodynamics become vol. energetics and the efficiency of processes in which no temp. drop is involved can be calcd. $A = \frac{1.97}{[(c_p/c_v)-1]} \left(\frac{1}{\delta_1} - \frac{1}{\delta_2} \right)$ or $\frac{1}{[(c_p/c_v)-1]} (KV_1 - KV_2)$; δ_1 and δ_2 represent the coeffs. of expansion and K represents the cohesion pressure. Thus 8750 cal./g. is obtained for the combustion of gas coal (mol. wt. 36, $c_p - c_v = 1.97$), 72

kg-cal for the combustion of H_2 and 2 v for the e m f of the H-O element. v. D gives a vol-energetic formulation of the 2nd fundamental law of thermodynamics

G SCHWOCH

New methods of establishing thermodynamic diagrams. J H CONLYN *Chaleur et ind* 10, 251-8(1929)—Mathematical *Ibid* 313-17—(1) The construction of the entropic curve for liquids, (2) the Mollier diagram and (3) variation of two parameters of the isochore as a function of the vol (for CO_2) are considered S L B E

Thermodynamics of mixtures. V FISCHER *Z Physik* 66, 269-79(1930)—The relations between osmotic pressure, absorption coeff and concn of C_2H_2 in Me_2CO , Me_2CO-H_2O , $Me_2CO-EtOH$ and $Me_2CO-MeOH$ are derived math and compared with the exptl results of Siller (Versuch über gelöster acetylen, Technische Studien Heft 5, Berlin Oldenburg, 1914)

GREGG M EVANS

Electric moments of organic molecules (HASSEL, NAESHAGEN) 10. Electric moments of some organic molecules in $C_{11}H_4$ solution (EIDE, HASSEL) 10.

Mitteilungen aus den Forschungsanstalten des G H H-Konzerns (*New journal*) Published by the Abt Konzern der Gutehoffnungshütte, Aktienverein für Bergbau und Hüttenbetrieb, Nürnberg, Germany. Can be obtained from V D I-Verlag G m b H, Dorotheenstr 40, Berlin N W 7 Vol 1, No 1 appeared Sept, 1930

BAVINK, BERNHARD Ergebnisse und Probleme der Naturwissenschaften. Eine Einführung in die heutige Naturphilosophie. 4th ed., revised and enlarged Leipzig S Hirzel 616 pp M 23, linen, M 25

CHAMOT, E M, and MASON, C W Handbook of Chemical Microscopy. Vol. II. New York John Wiley & Sons, Inc 401 pp \$4.50 Cf C A 24, 545

FLETCHER, GUSTAV L, SMITH, HERBERT O, and HARROW, BENJAMIN Laboratory Manual for Beginning Chemistry. 187 pp 76 cents Teachers' Manual and Key for "Beginning Chemistry" (title given in C A 23, 3151) 180 pp \$1.20 New York. Am Book Co. Reviewed in *J Chem Education* 8, 1008-9(1931)

HARVEY GIBSON, R. J Two Thousand Years of Science. 2nd ed., revised and enlarged by A W Titherley New York The Macmillan Co 503 pp \$3.50

J. C. Poggendorff's biographisch-literarisches Handwörterbuch für Mathematik, Astronomie, Physik, Chemie und verwandte Wissenschaftsgebiete. Band V. 1904 bis 1922. Berlin Verlag Chemie, G m b H 1423 pp M 142; half-calf, M 152

KOHLSCHEUTTER, H W Smith-Habers praktische Übungen zur Einführung in die Chemie. Karlsruhe B G Braun 140 pp M 5

KOPP, HERMANN Geschichte der Chemie. Reprint Four volumes bound in 2 buckram bindings Berlin Verlag Chemie G m b H Subscription price, M. 64, later, M. 80

LERBERGHE, VAN Calcul des affinités physico-chimiques. Paris Gauthier-Villars et Cie. 74 pp F 15

LORENTZ, H. A.: Lectures on Theoretical Physics. Vol. II. 410 pp \$7. Vol. III. About 21s., net. Translated by L Silberstein and A P H Travels London MacMillan and Co, Ltd

PHILBRICK, F. A.: Problems in Practical Physical Chemistry. London and Toronto. J M Dent and Sons, Ltd 146 pp 3s 6d

SCHAEFER, CLEMENS Einführung in die theoretische Physik. Band I. 3rd ed., revised and enlarged. Berlin Walter de Gruyter & Co 991 pp M 45, bound, M. 48 For Band II, cf C. A. 24, 546

SEEGERT, B.: Mikrophotographie. Berlin Union Deutsche Verlagsgesellschaft 56 pp M 6

WINKLER, L W.: Ausgewählte Untersuchungsverfahren für das chemische Laboratorium. Stuttgart: F. Enke About 176 pp About M 18, linen M 20.

3—SUBATOMIC PHENOMENA AND RADIOCHEMISTRY

W. ALBERT NOYES, JR

Recent advances in science: physics. L F BATES *Science Progress* 25, 570-6 (1931); cf. C. A. 25, 2357. JOSEPH S HEPBURN

Some simpler aspects of the corpuscular theory of light and the wave theory of matter. GILBERT D WEST *Science Progress* 25, 622-6(1931)—An attempt to harmonize the 2 theories JOSEPH S HEPBURN

secondary phenomena, particularly the reflection of electrons. For I a max. for the formation of ions occurred at 2-4 v., with a probability of 10^{-4} . Neg. ion formation in Hg vapor and A was shown only when electron speeds approached the ionization potential. In these cases neg. ions are formed probably from excited atoms. The probability curve for I reached a min. at 0.5 v. L. P. HALL.

Some new aspects of radioactivity. C. D. ELLIS. *Science Progress* 25, 607-21 (1931).—A review. JOSEPH S. HIEFELMAN.

A spinthariscopes for measuring radioactivity. H. A. MANGAN AND HERMAN SCHLEUBER. *Trans. Electrochem. Soc.* 59 (preprint) 7 pp (1931).—The spinthariscopes is serviceable in rapidly establishing the fact of radioactivity in living persons, but in its present form it is quite inferior to the electroscope for accurate estimation of small quantities of radioactive emanations. C. G. F.

The column ionization of single α -particles. K. DREIBER. *Physik Z.* 32, 181-2 (1931), cf. *C. A.* 22, 4365.—The degree of saturation of ionization of single α particles in air and CO_2 was investigated with varying angles between the direction of the incident beam and the electric field. A strong Po prepn was used giving an av. of 26 α particles per hr. At 15° and 1 atm., the const. K_0 in the Geiger formula is found to be for air, 6.26×10^4 , for CO_2 , $8.31 \times 10^4 \text{ cm}^{-1/2}$. The const. Z (atoms decomposed by 1 g. Ra per sec.) is 3.71×10^{11} . G. M. MURPHY.

The ionization of individual α - and H-rays at the end of their range. EWALD A. W. SCHMIDT AND GEORG STETTER. *Sitzb. Akad. Wiss. Wien, Abt. IIa*, 139, 123-39 (1930), cf. *C. A.* 24, 1279.—The tube electrometer was used to measure the ionization of many individual rays at various absorptions. α -Rays from Po prepn and H-rays (at fragments), obtained by α rays from Po prepn impinging on paraffin sheets, were studied. A special valuation method succeeded in approx. fixing the sp. ionization of a single α and a single H-ray at the end of their ranges as 20,000 and 5400 elementary quanta per cm^2 , resp., or a ratio of 4 to 1. By use of the tube electrometer this makes possible definite recognition of α rays of small range from H-rays. This definite separation should prove of value in investigations of α disintegrations. OGDEN E. SHEPPARD.

Recent progress in the study of the magnetic spectrum of α -rays. SALOMON ROSENBLYUM. *J. phys. radium* [7], 1, 439-44 (1930), cf. *C. A.* 24, 4457.—Resolution of α -rays from Th C was effected in a powerful magnetic field. Plates are shown for a direct method and a focal method. Relative velocities of 4 groups of rays from Th C referred to the strongest ray were measured as 1.0034, 0.975, 0.902 and 0.904, resp., and several other groups were detected, showing that particles are not all of the same velocity. The energies of these groups differed from the energies of the main Th C group by 40.6, 287, 442 and 441 kv. electrons. Certain groups of γ -rays possessed energies of 40.8 and 279 kv. electrons. L. P. HALL.

Investigations with the tube electrometer on the α -reflection and the shattering effect on light elements. EWALD A. W. SCHMIDT AND GEORG STETTER. *Sitzb. Akad. Wiss. Wien, Abt. IIa*, 139, 139-50 (1930).—The secondary radiations produced by α -particles from Po impinging on the light elements B, Be, C and N were investigated by the use of the tube electrometer in an attempt to determine whether at disintegrations took place. The measurements were made in an angle range of 135° and by comparison with analogous experiments using a thin Au foil it was determined whether a surplus of reflected α -particles was present. The values for Be and N are in good agreement with the theoretical but for B and C the experiments, with α -particles of 3.5-cm. residual range, show 4 to 4.5 times the theoretical no. of secondary α -rays. With α -particles of 2.3-cm. residual range C does not show this surplus of α -rays. With a special arrangement of the apparatus attempts were made to determine whether this excess of α -particles was due to anomalous scattering or to α decomposition but a definite decision is not yet possible. In the 90° angle range C showed the same particle no. as in the 135° angle range, which tends to disprove anomalous scattering. With α -particles of 2.3 cm. range B still shows clearly at decomposition. OGDEN E. SHEPPARD.

Application of Chamie's photographic method to the determination of the chemical behavior of polonium. MARILETTA BLAU AND ELISABETH RONA. *Sitzb. Akad. Wiss. Wien, Abt. IIa*, 139, 275-9 (1930), cf. *C. A.* 24, 1279.—Highly purified solutions of Po were mixed with the reagents considered and the mixture was then placed in a small closed cell in contact with the photographic plate and the plate developed. If there is uniform darkening of the plate the Po is considered to be in solution equally distributed but if there are dark spots on a lighter background it is considered that slightly soluble Po compounds are formed which collect in groups. The experiments indicate that N chromic acid and N HNO_3 do not repress all hydrolysis of Po salts but that N HCl does; KOH partly dissolves the hydrolysis products; glycerol tends to suppress hydrolysis, H_2S precipitates a Po compound.

which dissolves in 4 N HCl but not in $K_2S_2O_8$ reduces Po compds in H_2SO_4 soln but not in HCl. In hydrazine reduces Po compds in HCl soln., tartaric acid added to H_2SO_4 soln. produces more insol. compds. During electrolysis of Po compds in H_2SO_4 soln. 1-2 mm. dark spots appear in the photographs produced by anode soln. than in cathode soln. but if SO_2 absorbed in dark spots are seen in either. Further investigation of these electrolysis effects is planned in the hope of throwing light on the cathode migration of material during electrolysis.

OPPEL L. SHEPARD

Emanation apparatus. WINTNER L. LATITUDE 1 *Chem. Education* 8, 666-7 (1911).—An app. for the collection of emanation from a radioactive substance is described. The device is particularly useful in connection with lecture demonstrations and was designed to be used as a storage vessel as well as a working app. 111

The distribution of radon between liquid and solid phases of water and of benzene. FRANKS WITT. *Nat. Bur. Stand. Res. Abt. B* 130, 195-212 (1930).—Expts. showed that the quantity of Rn in the solid phase depends on the rate of freezing of H_2O or C_6H_6 solns. of Rn. Exhaustive expts. failed to show any simple relationship between rate of freezing and Rn content but did show that when $CaCl_2$ solns. are frozen very slowly Rn is not present in the solid phase. The phenomenon seems to be one of occlusion rather than soln. The fact that even very slow freezing did not reduce the Rn content of ice below 5-6% is probably due to the air bubbles enclosed in the ice. O. I. S.

Discoloration of glass and some minerals by β - and γ -rays. JOSEF HOFFMANN. *Nat. Bur. Stand. Res. Abt. B* 130, 213-231 (1930).—Many different glasses and mineral specimens were irradiated with β - and γ -rays from 610 mg. $RaCl_2$ until no further color change took place. Many samples were also irradiated with ultra violet rays from a quartz Hg vapor lamp. When heated most of the samples lost the color caused by irradiation but in some cases heating caused a change in color. For Mn free glasses the irradiation colors are yellow, reddish yellow, brown, reddish brown, gray and blackish. Alkali metal atoms are most commonly the cause of color, then Pb, Ba and Zn and certain heavy metal atoms follow. Only Pb glasses produce a reddish shade as a ground color. Pb glasses are recognizable because they show fluorescence under ultra violet light. Glasses relatively unstable because of high basicity, heavy metal oxide components, several different acid components, etc., frequently show first a gray color, then a brown and finally a blackish color. The Mn glasses present a unique problem which will be considered in a later paper. Violet is a common final color with Mn glasses.

OPPEL L. SHEPARD

Radium-bearing minerals from Great Bear Lake, Northwest Territories. HUGH S. SPENCE. Can. Dept. Mines, Mines Branch, *Memorandum Series* No. 43, 4 pp (1931).—Analyses of 2 pitchblende samples showed U_3O_8 56.91, 63.04, Th 0.1, U_2O_5 trace, U_4O_9 0.70, 0.66, Nitrate Co 0.13, 0.10, PbO 12.00, 12.13, Bi 0.18, trace, As 0.15, 0.14, Fe 1.01, 0.77, S 2.08, 0.90%, Ag 1.41, 1.72, Ant 0.19 oz per ton, trace. U_3O_8 content of 8 samples was 27.88-63.90%, equivalent to 70.79-213 mg per ton. Quartz was the principal impurity. Pb was present up to 11.25%. Flotation tests showed only little improvement in grade with a loss of 7.51% of the U_3O_8 .

ALDEN H. EMERY

The distribution of space-potential in high-frequency glow discharge. D. BANERJEE AND R. GANGULI. *Phil. Mag.* 11, 410-22 (1931), cf. *C. A.* 24, 5612. A. F.

The regularities of x-ray spark spectra. SAKAKI ICHI. *Science Repts. Tôhoku Imp. Univ.* 1st series 19, 551-8 (1930).—The values of $\Delta\sqrt{v/R}$ and $\Delta v/R$ between the so-called spark lines in the x-ray spectra and their main line have been studied as functions of the at. no., and it is found that $\Delta\sqrt{v/R}$ is nearly const. and $\Delta v/R$ varies linearly with the at. no. By these regularities it was possible to correlate certain spark lines found by several investigators.

C. J. HUMPHREYS

The fine structure of the Moseley curves for the M -absorption edges in the heavier elements. SAKAKI ICHI. *Science Repts. Tôhoku Imp. Univ.* 1st series, 19, 651-2 (1930), cf. following abstr.—Discontinuities in the Moseley curves for the L -absorption edges and for some emission lines in the L - and M series at $Pt(78)$ and near $Th(90)$ were reported in a previous paper. Similar discontinuities have been confirmed for the M absorption edges by the method of the modified Moseley curves.

C. J. HUMPHREYS

The precise measurements of the L -group of the x-rays in the heavy elements. SAKAKI ICHI. *Science Repts. Tôhoku Imp. Univ.* 1st series, 19, 559-639 (1930).—Complete results are presented of the measurements of the wave lengths of the x-ray lines in the L -group for the elements, $Ta(73)$ to $U(92)$. The wave lengths of the stronger lines determined by Firman and Wennerlöf were used as standards. Siegbahn's tube spectrometer was used throughout the expts. The differences in the values of v/R for the energy levels, $[L_1-L_2]$, $[L_2-L_{III}]$ and $[L_1-L_{III}]$ have been calculated from the data given by the emission lines

and are compared with the values obtained from the measurements of the absorption edges. The correlation of the exptl data with the process of atom building has been attempted by the method of the modified Moseley curves, which have been constructed for the elements studied. Regularities in the orders of the intensities as a function of the at. no. are discussed. C. J. HUMPHREYS

The fine structure of the Moseley curves for the energy levels in the lower elements. SAKAE IDEI. *Science Repts Tôhoku Imp Univ.* 1st series, 19, 641-9 (1930), cf preceding abstr.—The fine structure of the Moseley curves for the energy levels in elements of low at. no. has been studied by a simple method proposed by I., according to which the curve is assumed to be a straight line, $\sqrt{\nu}/R = a + bZ$, after which the differences between the actual values of $\sqrt{\nu}/R$ and corresponding points on this line are plotted against Z . Such a difference curve is called a modified Moseley curve. Many discontinuities in the curves studied corresponding to the stepwise process of atom building are brought out in a striking manner. Some considerations are put forward on the screening effect of other electrons. C. J. HUMPHREYS

The fine structure of the K-absorption spectrum of selenium. SAKAE IDEI. *Science Repts Tôhoku Imp Univ.* 1st series, 19, 653-8 (1930).—Measurements of the fine structure of the K-absorption edge of Se are reported. The effect of chem combination on the principal K-absorption limit has been found to be negligibly small. Suggestions which have been made as to the possible origin of such structures will require data from a larger no. of elements for confirmation. C. J. HUMPHREYS

Theory and practice of x-ray analysis. WILLIAM H. BARNES. *Gen Chem Met* 15, 67-70 (1931).—A short review with brief mention of some of the fundamental conceptions on which the subject is based and an outline of a few of the fields of application. The production, nature of x-radiation and phys chem and biol effects thereof are discussed. Applications to spectroscopy and crystallography are noted. W. H. B.

Construction and operation of a simple x-ray spectrograph. LYMAN J. WOOD. *J Chem Education* 8, 952-63 (1931). E. H.

Quantitative chemical analysis by means of the absorption of x-rays. N. H. MOXNES. *Z physik Chem., Aht A* 152, 380-403 (1931), cf C. A. 24, 548.—An analytical method based on the absorption of x-ray line spectra was developed. A photographic-photometric method for measuring the intensity of x rays was described. When it was applied to the analysis of Zn in ZnO and Ni in NiO, the max errors were 10 and 39%, resp. FRANK URBAN

Protective lead thickness in the German protective recommendations. J. H. VAN DER TUUK AND W. H. BOLDINGH. *Fortschr Gebiete Röntgenstrahlen* 41, 905 (1930).—The conclusion is drawn that the international recommendations regarding Pb thickness for x-ray protection are preferable to those of the German Röntgen Society. Dosage measurements were made to obtain the necessary data. H. E. SEEMANN

Protective lead thickness in the international and in the German protective recommendations. R. GLOCKER. *Fortschr Gebiete Röntgenstrahlen* 41, 907 (1930).—G. refers to the work of van der Tuuk and Boldingh (preceding abstr.), and points out the difference in principle used in deciding upon the international and the German recommendations. A decrease in thickness from 5 mm to 4.5 mm for 200-220 kv appears desirable. H. E. SEEMANN

The electronic energy levels of the elements, with special reference to their connection with the sizes and electronic states of atoms in metallic crystals. WILLIAM HUME-ROTHERY. *Phil Mag* 11, 649-78 (1931), cf C. A. 24, 5559.—The size of the atom in a crystal is detd principally by the outermost complete shell of electrons of the at. core or ion and not by the valency electrons. When the interat distances vary as $1/Z^2$ at the beginning of a period, the electronic energy levels vary linearly with Z^2 for the N_1 electrons of the outermost group of electrons of the at. core. This rule was tested for the various groups by plots of ν/R against Z^2 or for the higher powers $\log(\nu/R)$ against $\log Z$. The slope of the line when the interat. distance is plotted against Z is $1/2$ that of the slope of the energy level line and with the opposite sign. From this the no. of electrons in the outer shell may be deduced. The rule is applied to various elements. ARTHUR FLEISCHER

Excitation of the visible spectrum of helium. J. S. TOWNSEND AND F. LLEWELLYN JONES. *Phil. Mag.* 11, 679-85 (1931).—The intensity of red light for a given current increased 27 to 1 when the pressure was decreased from 10.4 to 2.3, for violet light the intensity increased 4.3 times. The calcd. mean energy of the electrons is 3-4 v. This is not in agreement with the properties of the Bohr model of He, which requires 20 v. for radiation emission. ARTHUR FLEISCHER

The magneto-optical dispersion of organic liquids in the ultra-violet region of the spectrum. II. The magneto-optical dispersion of methyl propionate, ethyl propionate and ethyl formate. J. W. JENKINS AND F. J. EVANS. *Phil Mag* 11, 377-106(1931); cf. C. A. 24, 141. The constants λ_1 and $\lambda_2 \times 10^{-2}$, resp., for the equation given in C. A. 24, 141 satisfactorily explain the experimental data at 19° for methyl propionate 0.1066, 4.771; ethyl propionate 0.1061, 4.72. The values of e/m for the two liquids calculated were 0.96 0.99 and 0.996 resp. The ordinary dispersion may be expressed by the equation $n^2 - 1 = b_0 + \beta/(\lambda^2 - \lambda_1^2)$ with the constant b_0 0.8716, 0.81 0.82, and constant $b_1 \times 10^{-1}$ 8.68, 8.93 and 8.30, resp., for the 2 compounds.

ARTHUR LEITCH

The optical excitation function of helium. WALTER C. MICHELS. *Phys Rev* 36, 94 (1931).

BERNARD LEWIS

Vacuum spark spectra to 40 Å. U. the spectra of Be III, Be IV, B IV, B V and C V. R. S. T. JONES. *Nature* 127, 403-6(1931), cf. C. A. 24, 2049.—The series of II and He-like spectra which was previously traced to Be IV, has now been completed with B IV, B V and C V and the limit of optical spectra brought down to 40.28 Å. U. The vacuum spark was produced by a capacity of 0.3 μ f charged to 60,000 v. Half an hr with about 30 sparks a min. was sufficient for an exposure. The plates were taken with the same metal grating as was used before, but now set up in a new spectrograph at a glancing angle of 74°. Tabulated results for the 2 series of isoelectronic spectra are reproduced below.

	Be IV		B V	
	λ	ν	λ	ν
1'S - 2'P	1,317.084	75 925	2,058.247	48 585
1'S - 3'P	1,550.903	64 063		

	Be III		B IV		C V	
	λ	ν	λ	ν	λ	ν
1'S - 2'P	100.25	997,500	60.31	1,658,100	40.28	2,482,600
3'P	88.30	1,132,500	52.68	1,898,300		
4'P	84.75	1,179,000				
5'P	83.19	1,202,100				
6'P	82.37	1,214,000				

The ionizing potentials, computed on the assumption that the quantum defect, $n - n^* = 0.013 = 0.001$, calculated from the series of Be III, is constant, are in good agreement with the quantum theoretical values given by Hylleraas.

C. J. HUMPHREYS

Note on wave lengths in the vacuum copper arc. A. G. SHENSTONE. *Phys Rev* 36, 602-3(1930).—For the measurement of Cu II low level spark lines, spark spectra lines are satisfactory standards. However, for high level lines an error of about 27 wave nos. is introduced by the same procedure.

BERNARD LEWIS

Fluorescence of mercury vapor under atomic and molecular absorption. HENRYK NIEMCOWICZANSKI. *Nature* 127, 406(1931), cf. Rayleigh, C. A. 25, 2051.—Lord Rayleigh has recently pointed out that the fluorescence of Hg vapor is due to both at. and mol. absorption. This conclusion is supported by experiments on the influence of the magnetic field on this phenomenon. The fluorescence was excited by radiation of 2537 Å. U. The tube with the Hg vapor was placed between the poles of an electromagnet and the intensity of the fluorescent radiation observed at different field strengths. The variation of the intensity of the green fluorescence was found to be similar to that of the resonance radiation investigated by M. Scheem. It is concluded, therefore, that, in the excitation of the fluorescence by 2537, the at. absorption forms the first necessary step.

C. J. HUMPHREYS

Intensity, broadening, extinction and secondary continua of the mercury resonance line on addition of rare gases. PAUL KUNZE. *Ann Physik* 8, 500-20(1931).—The effects of He, Ne and Ar on the emission and absorption of Hg 2537 have been further investigated. Monochromatic light is passed into a cell, connected directly with this resonance cell is an absorption cell containing the same mixt. of gases. The photoelectric cell used to measure the resonance is on a line with the resonance and absorption cells. All measurements were made at 21.5°. The pressures of the rare gases varied from 0 to 1 atm. Absorption is greatly decreased by even small quantities of the gases, with increasing pressure the 3 curves (absorption vs. pressure) become almost coincident and are almost parallel to the pressure axis, although still decreasing. The sp. broadening according to the Lorenz theory should be 1.1 Å for He, Ne and Ar, explicitly it is 18, 12 and 19 mμ. A. U. resp. The extinction of the resonance intensity decreases in the order He,

Ne and Ar; for He it is almost 0, for Ne small and for Ar marked. A discussion is given. WILLIAM F. VAUGHAN

The presence of neutral oxygen in the gaseous nebulae. I. S. BOWEN *Phys. Rev.* 36, 1400 (1930). HERNARD LEWIS

The two forms of hydrogen. T. J. DEBIA CRUZ *Magyar Chem. Folyóirat* 36, 1-10 (1930) —A theoretical discussion of para- and ortho-H. S. S. DE LINIA

The ortho-para-hydrogen conversion at surfaces. HERBERT S. TAYLOR AND A. SHURMAN. *J. Am. Chem. Soc.* 53, 1011-5 (1931) —Neither with an active ZnO nor with the more active hydrogenation agent Zn-Cr oxide is there any para-H formed at liquid air temps. even after 14 hrs. contact. In the temp. region where, on these surfaces, activated adsorption of H sets in ($T > 0^\circ$) there is an immediate activity in the re-conversion of the 50% para-H to the normal 1:3 mixt. The relative activities in this respect are exactly parallel with the influence of compn. on the velocity of activated adsorption. C. J. WINE

The Stark effect in band spectra. W. G. FENNIE *Phil. Mag.* 11, 102 (1931) Math. Appreciable Stark effect may be expected when an elec. field is applied to a radiating mol. system when for dist. mol. $\mu \neq 0$ and there is a non-vanishing permanent moment. The splitting will be quadratic at first, then linear when the energy shifts become large compared with natural Δ doubling intervals. A similar result is obtained for polyat. mols. with 2 nearly equal moments of inertia. For polyat. mols. with 3 appreciably different moments of inertia only the lowest quadratic states will give a measurable splitting. ARTHUR LEWIS

The Raman effect in acetylene. I. SIECKE *Atti. accad. Lincei* 12, 226-8 (1930) The Raman effect in gases has not been studied very extensively as in many cases the ultra-violet rays usually used (from Hg) have a chem. action on the gases. S. therefore utilizes some of the visible rays from the Hg spectrum in studying the action on C_2H_2 . Five sharp lines and one faint line were obtained in the spectrograph. The 5 sharp lines were at $\nu = 25,117, 25,377, 25,318, 22,725$ and $20,950 \text{ cm}^{-1}$, excited by the 3650, 4016 and 4154 Å U. lines of Hg. The difference in displacement in various spectrographs was $\Delta\nu = 1079 \pm 0 \text{ cm}^{-1}$. The 6th line is of about 21910 cm^{-1} . The frequency 1079 cm^{-1} has also been observed in the infra-red absorption and is one of the fundamental frequencies of C_2H_2 . A. W. CONNELL

Raman spectrum of hydrogen peroxide. S. VENKATSWARAN *Nature* 127, 103 (1931) —The Raman spectrum of H_2O_2 (Merck's perhydrol 30% soln. in water) has been photographed. The prominent feature is the presence of a sharp Raman frequency of 875 accompanied by a weak component at 901 cm^{-1} . The low values of the frequencies suggests that they arise from O atoms, and as such may be compared with the O oscillation frequency 1552 obtained by McLennan (*C. A.* 24, 1291). The large difference between the 2 values is indicative of the fact that the O atoms in H_2O and O_2 mols. resp. are bound differently. Besides the above 2 frequencies there appear other bands which are less intense and sharp, the origin of which is under investigation. C. J. H.

The electrolytic dissociation of nitric acid as revealed by its infra-red absorption spectrum. I. L. KINSLEY AND J. W. PETER *Phys. Rev.* 35, 603-4 (1930) —In solns. of HNO_3 of various concns. and of the fuming acid only the water bands in absorption were found. They were sharp and shifted slightly toward shorter wave lengths. In the fuming acid new strong maxima appear which are attributed to the undissociated HNO_3 mol. In the vapor a group is found which is believed to be the same as that for the liquid. Each band, however, is shifted to shorter wave lengths. H. L.

Photochemical decomposition of nitrogen trichloride. I. JOHN G. A. CRIPPIEN AND EDWARD G. W. NARRIS *Proc. Roy. Soc. (London)* A130, 691-699 (1931) —The decompn. of NCl_3 in Cl_2 alone is sensitized by Cl_2 and the net reaction is $2NCl_3 \rightarrow N_2 + 3Cl_2$. The reaction is strictly of 0 and order with respect to NCl_3 until the decompn. is practically complete; a semi-explosive reaction then occurs. The velocity is directly proportional to the light absorbed by the Cl_2 . Quantum efficiencies were measured in 4360 Å U. light and 4050 Å U. light; pressures of NCl_3 ranged from 0.116 to 2.48 mm; Cl_2 pressures varied from 22.3 to 700 mm. The quantum efficiencies are high (between 17 and 5) at low concns. of Cl_2 (between 22 and 100 mm) and approach a limiting value in the neighborhood of 3 at the highest pressures of Cl_2 . The results are reproduced by the equation $\gamma = (1/0.0014)(1/psn) + 2.5$. No decompn. occurs in green or yellow light. The app. is described. A short discussion is given, G and N, postulate an initiation of the reaction by Cl atoms; it then proceeds by way of reaction chains of short length. WILLIAM F. VAUGHAN

The oxidation of sulfur dioxide in ultra-violet light. G. KORNFIELD AND R. WERNMANN. *Z. Elektrochem.* 36, 781-91 (1930) —The SO_2 and O_2 mixt. was exposed in small

quartz vessels. The SO_2 content was determined iodometrically. The illumination was conducted with light of the following wave lengths: 3130, 3020 to 2970, 2800, 2537, 2300 to 2130 Å. C. with the use of a monochromator. At 2537 Å. U the conversion was less than 0.01 mol. per quantum. For the same wave length the quantum efficiency, ϕ , for conversion is dependent upon the concn. of SO_2 , decreasing as the concn. of SO_2 increases. At 3130 Å. U the decrease is greater than at 2970 Å. U. The reaction velocities are independent of the concn. of SO_2 and O_2 . Any temp. influence on the velocity between room temp. and 65° was not ascertained. Corresponding quantum efficiencies are for (1940 Å. I = 100), (2650 Å. U = 0.71), and (2300 Å. U = 0.53). A partial discussion of the reaction mechanism was based on the disson. of Q since the reaction limit is concordant with the range of the Q disson. The absorption spectrum of SO_2 has 2 separate regions between which the one reaction lies. Conclusion: The mechanism of the decomposition reaction in the long wave length region must be entirely different from that in the short wave length region in view of energy considerations.

W. VASCHOW

Experiments on the luminescence of zinc sulfide and diamond under the influence of radioactive radiation. HERTA KARNIK. *Wiss. Abt. Hist. Mus., Abt. IIa, 119, 303-19 (1930)*.—The decrease in luminescence of a ZnS screen under the effect of pure α radiation from Po was measured. The results are compared with those of Marsden (C. A. 5, 242). A filter was found which stops the β bands in the emitted light. The individual variations of these bands was studied. Another sample of ZnS of smaller scintillating ability showed a smaller decrease in luminescence. The Rutherford formula was found not to fit the data when applied to the individual bands. A value for γ radiation was found. This was shown not to be caused by regeneration in the dark. Expts. were made on the effect of α - and γ radiations on a sample of diamond. The results were similar to those obtained with ZnS. A description is given of the method by which the ratio of luminescence intensity to particle density was controlled.

E. J. ROSENBACH

Band width and temperature dependence of the emission bands of phosphorescent alkali halides. FRANZ URBACH. *Wiss. Abt. Hist. Mus., Abt. IIa, 119, 249-52 (1930)*, cf. C. A. 24, 1280.—U. has derived a theoretical equation which requires that (1) for bands of the same width, the displacement with temp. is the greater for those bands which show the larger deviation from a linear relationship between the band width and \sqrt{T} . (2) for bands of equal deviation from the \sqrt{T} law, the broader bands show the greater displacement. These predictions have been checked against the data of v. Meyer (C. A. 24, 4253). Within the limits of error the predictions are fulfilled.

E. J. ROSENBACH

Luminescence of the alkali halides. I. Introduction and orientating observations. II. Methods of measurement—first results—theory of thermoluminescence. FRANZ URBACH. *Wiss. Abt. Hist. Mus., Abt. IIa, 119, 353-62, 363-72 (1930)*.—For a series of alkali halides, NaCl, KBr, RbBr, RbCl, CsCl and natural rock salt andylvine, U. made preliminary visual observations of the bands emitted around 5000-3000 Å. U., during and after exposure to Ra emanations. The phosphorescence on heating and illumination and the thermoluminescence were observed. Rock salt showed a blue phosphorescence, while the artificial NaCl showed a yellow, which could be changed to blue on melting or applying pressure. The blue of rock salt could not be changed to yellow. KCl gave a sec. intensity max. of fluorescence when illuminated again after a pause of a few sec. Forylvine it was found that the total amt. of fluorescence was greater under illumination than from heating. An estimate of the proportion of active centers leads to the order of magnitude 10^{-8} . Exptl. details are given for quant. measurements, made by means of photoelec. cells. The existence of the recuperation effect was confirmed. Curves are presented showing the variation of thermoluminescence with time. A theory of thermoluminescence is developed, based on the assumption that an activated center must pass a threshold energy level before it can return to an unactivated state. III. Measurements of thermoluminescence, fluorescent-center hypothesis. FRANZ URBACH and GEORG SCHWABE. *Ibid.* 483-95.—Errors in the measurement of the thermoluminescence of the alkali halides are discussed critically. Curves are presented for data obtained by varying the time of irradiation with Ra, the color of the illumination, the state of subdivision, and pressure applied before and after irradiation. Reproducible results could be obtained only by tempering the crystals at 200° for 24 hr. The results are interpreted with the aid of the luminescent-center hypothesis, modified by ascribing a finite life to the luminescent centers which is dependent on the temp.

E. J. R.

Luminescence analysis. III. The alkaline earth group and quantitative knowledge of luminescence. FRIEDR. BEUTEL and ARTH. KUTZELNIGG. *Monatsh.* 57,

15-0(1931), cf C A 25, 2579—CaO, Ca(OH)₂, CaCO₃, CaSO₄ and the corresponding compds of Ba and Sr were investigated with a photometer through red, green and blue filters. The results are given tabularly. GREGG M EVANS

Experiments on radiothermoluminescence. FRANCES G WICK *Sitzb Akad Wiss Wien, Abt IIa*, 139, 497 508(1930), *J Optical Soc Am* 21, 223-31(1931)—Quant observations were made of the thermoluminescence, after irradiation with Ra of synthetic CaSO₄ with added Mn and of samples of fluorite. The photoelec method of Urbach was used, exptl details are given. The CaSO₄ used showed a single intensity max which was lowered by pressure applied either before or after Ra irradiation. Fluorite showed a series of max which are ascribed to the rare earths on the basis of a spectroscopic examn. These max also were lowered or destroyed by pressure, while a blue band was relatively intensified by pressure applied before irradiation. The natural thermoluminescence of the fluorite samples was examd. The effect of Ra irradiation on the substances investigated is similar to that obtained with Rontgen rays. E J R

Determination of ultra-violet-ray content of sunlight. TORATARO HANZAWA *J Agr Chem Soc Japan* 6, 1093-102(1930)—A chem actinometer using NiI₂ molybdate in dil HCl with small quantities of EtOH is suggested. Y KIHARA

Photochemical action of intermittent and complex light. M PADOA AND N VITA *Z wiss Phot* 28, 153-67(1930)—It was found experimentally that under certain conditions the photochem reaction of HI oxidation with filtered light gives a higher effect than with white light. It was shown that Berthoud's law holds for HI oxidation in white light and that, for monochromatic light, for each wave length there is a different ratio between intensity and effect. A P H TRIVELLI

Photovoltaic studies on metal and oxide electrodes in distilled water and in dilute solutions. B LÁNYI AND E TURESZ *Magyar Chem Folyóirat* 35, 124-39, 145-51 (1929)—Ag, Cu, Ag₂O and CuO electrodes were studied in distd water and in dil solns. Expts were carried out both in the absence of foreign gases and with various gases bubbled through the solns and over the electrodes. Pure metals showed effects up to 1 mv, while metals with oxidized surfaces showed effects up to 200 mv. The effects were the same in distd water as in dilute solns of electrolytes. H, N and O were without effect. An explanation is offered based on a photochem reaction between the electrode and the soln. This reaction is probably reversible in the dark, since the photovoltaic potential gradually becomes zero after illumination ceases. S S DE FINÁLY

Photophoresis and the influence upon it of electric and magnetic fields. FELIX FÜRENFART *Phil Mag* 11, 140-6(1931); cf C A 22, 1723—With transverse magnetophotophoresis, the movement of the particles is reversed on reversing the magnetic field. This reversal does not occur in longitudinal magnetophotophoresis. With a transverse elec. field, particles on illumination may move in either direction. Te, Sb, and I particles move at low fields strengths, while Ni, Fe, Se and Bi require high strengths. No effect is observed in an alternating elec. field. The effect is capable of quant measurements. ARTHUR FLEISCHER

Azo-chromophore (spectra of about 20 dyes of the *o*- and *p*-amino- and -hydroxy-azobenzene types) (BLUMBERG) 25. The magneton numbers of the Fe group (DE HAAS, GORTER) 2. The kinetics of surface processes on crystal lattices (MURR) 2.

ALLEN, H S *The Quantum and Its Interpretation*. London Methuen & Co 214 pp 12s 6d Reviewed in *Trans Faraday Soc* 27, 191(1931)

DHAR, N R *The Chemical Action of Light*. London Blackie and Son 512 pp 25s, net

Ergebnisse der technischen Röntgenkunde. Edited by J EGGERT AND E SCHIEBOLD. Band II. Fortschritte der Röntgenforschung in Methode und Anwendung. Die Heidelberger Röntgengtagung, 1930. Edited by F KÖRBER AND E. SCHIEBOLD. Leipzig Akad Verlag 326 pp M 36 60, bound, M 38

SCHLEEDER, ARTHUR, AND SCHNEIDER, ERICH *Röntgenspektroskopie und Kristallstrukturanalyse*. Band II. Berlin Walter de Gruyter & Co M 20, linen, M. 22 Cf C A 23, 3167

SHAPLEV, HARLOW *Flights from Chaos: A Survey of Material Systems from Atoms to Galaxies*. New York Whittlesey House, 168 pp \$2 50 Reviewed in *J. Franklin Inst* 211, 528(1931)

WIGNER, EUGEN. *Gruppentheorie und ihre Anwendung auf die Quantenmechanik der Atomspektren*. Brunswick Friedr Vieweg & Sohn Akt.-Ges. About 325 pp About M 20, half linen, about M 22

4-ELECTROCHEMISTRY

COLIN G. FINK

Advantages of electric melting furnaces. F. W. FRIESE. *Metallurgia* 2, No 10, 124 (1911) — Performance and comparative cost data are discussed, with particular reference to Cu alloy foundry work in Brazil. F. I. S.

Factors affecting the economics of industrial electric heating. N. R. STANSFEL. *Trans. 2nd World Power Conference* (Berlin) 1, 429-40 (1930) — Of metallurgical interest. ALDEN H. MERY.

Electrically heated annealing furnaces for the steel and other industries. J. C. WOODSON. *Iron Steel Eng.* 8, 122-8 (1931) — Annealing covers full annealing, process annealing, normalizing, patenting, spherizing, malicabilizing, graphitizing, etc. Brass annealing. Cu, Al, steel, castings and forgings, special and fabricated shaped, wire strip and sheet annealing, bright annealing in artificial atm., and future trends are all discussed. W. H. BOYNTON.

Electric heating insures quality production. H. J. FISCHBECK. *Elec. World* 97, 680-1 (1931) — Elec. heat is employed by the Pratt & Whitney Aircraft Corp. as a result of the following advantages: (1) high efficiency with attendant quality of work, (2) simplicity of application, (3) close automatic temp. control, (4) cleanliness and improvement of working conditions, (5) impossibility of overheating any part of the charge and (6) elimination of the combustion method with its attendant uncertainty of results and general fuel problems, plus reduced fire hazards. Carburizing, hardening and miscellaneous heat treating are handled. W. H. BOYNTON.

Difficult production problem solved by electric heat. L. J. LARSON. *Elec. World* 97, 723 (1931) — Asphalt-coated pipe requires careful control of coating temps. during the period of immersion and solidification. Roughly speaking, the temp. at which the compd. is applied depends inversely upon the thickness of coating desired.

Coking of the coating must be carefully avoided. A steel tank with concentric resistance heater frame warms the pipe compd. A large order of asphalt-coated pipe was handled in a vertical, electrically heated tank. The coating material melts at 102° and is applied at 204°. W. H. BOYNTON.

Economy of electric resistance furnaces from the standpoint of power industry and consumer. V. PASCHKE. *Elektricitätswirtschaft* 29, 474-9 (1930). E. I. S.

Electric resistance furnace. KNOORS. *Elektricitätswirtschaft* 29, 465-73 (1930) — The advantages of elec. furnaces for the power industry and the consumer are discussed, and the principle types are described. Their field of application is outlined and consumption data are given. E. I. S.

Electrodes of the third order. J. VELISEK. *Chem. Listy* 24, 443-7 (1930) — An electrode of the 3rd order consists of a metal surrounded by 2 insol. (depolarizer) salts which also sat the soln. The first depolarizer has a cation corresponding to the metal of the electrode, the potential depends only upon the concn. of the cation in the electrolyte. The electrodes of the third order are used practically to det. concns. of such metals (Ca, Ti) which decompose with H₂O and cannot be constructed as electrodes of the first order. The system Pb | PbSO₄, CaSO₄, Ca⁺⁺ is discussed. The system Hg | Hg₂(IO₃)₂, Zn(IO₃)₂, Zn⁺⁺ showed stable potentials only after 3 days. With freshly silvered electrodes, the system Ag | Ag₂C₂O₄, CaC₂O₄, Ca⁺⁺ became stable in a few min. The system Zn | ZnC₂O₄, CaC₂O₄, Ca⁺⁺ became const. after 20 min. and remained dependable, the Zn was amalgamated. The system Hg | Hg₂IO₃, TlI, Tl⁺⁺ is discussed. Computations of potentials of the various systems are given. F. M.

A Luther calcium electrode of the third order. J. VELISEK AND K. SVENCON. *Chem. Listy* 24, 467-71 (1930) — The electrode Pb | PbSO₄, CaSO₄, Ca⁺⁺ was constructed, it fulfils the conditions postulated by Luther for electrodes of the third order. The second depolarizer is CaSO₄, its soly. is too great to det. Ca ions accurately and yields as many Ca ions as a pure 0.01 N CaCl₂ soln. The potentials of individual electrodes (when compared to a standard Hg electrode) varied among themselves and fluctuated in the same series of measurements, although systematic changes in the potentials were not observed and indicate that all secondary reactions that Luther described progress so slowly that the potential becomes practically independent of the time. For more accurate measurements of Ca ions another electrode must be devised, in which the second depolarizer is less sol. FRANK MARESH.

Electrolytic refining of aluminum and aluminates in an aluminum chloride-sodium chloride melt. V. A. PLOTNIKOV, N. S. FORTUNATOV AND V. P. MASHOVITZ. *Z. Elektrochem.* 37, 83-8 (1931) — A general review of patents and a discussion of all the

hrs. at 120-4) and 0.2531 (44.9% theoretical) 0.2531 (44.9%), 0.3164 (56.2%), and 0.2768 g. (49.2%) of I obtained by hydrolysis was detd. by oxidizing to AcMe and colorimetrically estg. with *o*-nitrobenzaldehyde. This shows that III adds on to IV and hydrolysis of the ester formed yields I, the reaction indicating a possible origin of this alc. during electrolysis. In order to det. the possibility of step 2 the thermal and photochem. decompn. of butyryl peroxide (V) was studied. 0.3430 g. of 90.2% V heated in a bronze bomb in an oil bath at 271° gave 99.18% of the theoretical quantity of CO₂. The yield of hexane in another detn. was 62.6% of the theoretical. 0.1910 g. of 91.1% V in a quartz tube was exposed to a Hg vapor lamp for 22 hrs. and yielded 4.6% of the theoretical amt. of CO₂ but only traces of hexane were obtained, and a colored residue remained in the tube. In order to det. the possibility of step 3, weighed samples of V were dissolved in 10 cc. AcMe, 10 cc. of IV and 100 cc. water were added, the mixt. was shaken at const. temp. and samples were periodically analyzed to det. the extent of hydrolysis. The samples were cooled to 0°, dild. with AcMe and KI in AcOH, and titrated cold with Na₂S₂O₄ to det. per acid formed and then heated to 30-40° and again titrated to det. V. In one case 0.1591 g. V at 31-32° showed 4.6% hydrolysis after 60 min., 7.14% after 150 min., 8.44% after 200 min. and 9.33% after 300 min. At 17-18° 7.34% hydrolysis occurred in 300 min. In order to det. the possibility of reaction 4 the thermal and photochem. decompn. of butyric per acid (VI) was studied. 0.5450 g. of 87.2% VI in a bronze bomb was heated to 270° and 0.1775 g. CO₂ (88.4% theory) and 0.1768 g. III as propylene dibromide (VII) (92.1% theory) obtained. Four g. of VI after 1½-hrs. exposure to a Hg vapor lamp yielded only 0.150 g. VII and a little CO₂ and indications were that more of VI decompd. to form O₂, solid hydrocarbons and resins. OPEN E. SHEPARD

Principles of the hydrometallurgy and electrodeposition of the metals. T. P. CAMPBELL. *Colo. School of Mines Mag.* 20, No. 3, 13-15, 45, No. 4, 15-19, No. 5, 25-6 1930.—A detailed review. E. I. S.

The cuprous-cupric equilibrium in copper sulfate and perchlorate solutions and the anodic behavior of copper. C. HERVET. *J. Electrochem.* 37, 61-76 (1931).—A detailed and practical description of the previous studies made on the equl. $2\text{Cu}^{++} \rightleftharpoons \text{Cu}^{+} + \text{Cu}$. The expts. carried out in this study consisted in the analytical part of a) titration with K₂Cr₂O₇, (b) titration with Ce(SO₄)₂. Very pure Cu salts required in the expt. were prepd. The equl. was studied at 25° to 60°. In a molar CuSO₄ soln. there are 0.71 millimoles of Cu⁺ in equl. at 25° and 3.15 at 50°. In sulfate solns. the ratio $\sqrt{\text{Cu}^{++} \text{ molarity}} / \text{Cu}^{+} \text{ molarity}$ is independent of the salt concn. within wide limits. From the values obtained with sulfate soln. the reaction $\text{Cu}^{++} + \text{Cu} = 2\text{Cu}^{+}$ at temps. of 25-60° gives a value of 18,500 cal. and a calcd. normal potential of 0.525 v. for the reaction $\text{Cu} \rightarrow \text{Cu}^{+}$ at 25°. G. T. MORON.

A study of silver-plating solutions. B. EGERBERG AND N. PRONITZ. *Trans. Electrochem. Soc.* 59 (preprint) 19 pp. (1931).—A discussion is given of the effects of temp., c. d. and concns. of Ag, cyanide, carbonate and CS₂ on the reactivity, throwing power, and cathodic and anodic polarization in argentocyanide plating solns. Curves are shown. A comparison is made between solns. contg. Na salts and those containing K salts, and the advantages of the latter pointed out. Optimum plating conditions and compns. for the above solns. are suggested. The subject is treated mainly from the industrial standpoint, although theory is also briefly considered. C. G. F.

Electrodeposition of silver from sulfate, nitrate, fluoborate and fluoride solutions. EDWARD B. SANCAR. *Trans. Electrochem. Soc.* 59 (preprint) 21 pp. (1931).—The possibility of electrodepositing Ag in a form suitable for electroplating purposes from silver sulfate, nitrate, fluoborate and fluoride solns. has been studied. Owing to the simple-immersion deposition of Ag from these solns. upon Cu, Hg, brass and Ni, adherent deposits were obtained only upon Ag. Relatively compact deposits of Ag were obtained upon Ag deposited from an argento-cyanide soln., from Ag₂SO₄ solns. contg. boric acid or Fe₂(SO₄)₃, Ag₂O solns. contg. boric acid, silver fluoborate solns. and silver fluoride solns. However, these solns. showed a great tendency to yield treed deposits, and the conditions under which deposits free from trees could be obtained were limited. All the compact deposits were decidedly cryst., and while they might be of use in the electrorefining of Ag, as a means of obtaining dense Ag deposits instead of the loose crystals given by the nitrate soln., they offer no serious competition to the usual argento-cyanide soln. for the electroplating of Ag, unless considerable improvement can be made, for example, by the use of "addition agents". A further difficulty in the use of these baths for Ag deposition is the unsatisfactory corrosion of the anodes. C. G. F.

Electroplating on aluminum. H. K. WORK. *Metal Cleaning and Finishing* 2, 777-80 (1930).—See C. A. 22, 3844, 25, 2370. E. I. S.

Plating glass and clay. I WINKLER *Speerschall* 63, 221 (1930) In order clay or porcelain conductive, cover with graphite compound and heat to 200° or 400°

I: 1 S

Improvement in electroplating quality needed to meet competition. I B NICH *Automotive Ind* 63, 560-4 (1930) The quality of plating and future protective progress are discussed. Ni plating at low pH , in treatment of Cr plate, durability tests, control methods, Cr plating for resistance to abrasion and electrodeposition of Fe are also discussed

I: 1 S

The influence of foreign metals during nickel plating. B SIVIK *Chem Listy* 25, 123-4 (1931)—A Ni bath (H_2SO_4 30%, $Ni(NH_4)_2(SO_4)_2$ 70%) contg 1% $CuSO_4$ (0.25% Cu) gave a dark and stained deposit or showed dark spots with 5% $CuSO_4$ (1.26% Cu) Ni plating was not possible, with 10% $CuSO_4$ (2.53% Cu), the Cu was pptd as a powder. All of the Cu was easily removed upon a metallic cathode with a strong current. The addn of 0.5% $ZnSO_4$ (0.11% Zn) to the Ni bath retards the rate of Ni deposition, with 0.266% Zn stains form over the deposit with 0.45% Zn the bath becomes useless. The addn of 0.5% $FeSO_4$ (0.10% Fe), 0.20% Fe, and 0.403% Fe to the Ni plating bath showed no effect, with 0.601% Fe the bath plates well but gives a dark deposit, the addn of 0.02% Fe leaves imprinted areas especially at the point of support and the deposits are dark and irregular. Analysis of 4 brands of Ni showed Ni 96.7 (99.2) 0.54 2.01 and Cu 0.11 0.10%. Analysis of 2 brands of Ni $SO_4 \cdot 6H_2O$ showed traces of Cu and 0.23-0.15% Fe

I: 1 S

Further developments in low- pH nickel deposition. W M PHILLIPS *Trans Electrochem Soc* 59 (preprint) 4 pp (1931) cf *C A* 24, 4090—Further commercial work indicates that (1) deposits from baths of low pH furnish better protection than those from baths of high pH , (2) the hardness of deposits is not directly influenced by the pH . Ni is softer from baths operated at elevated temps and (3) pitting is no worse at low pH than it is at high pH

C G F

A note on the tarnishing of chromium-plated brass. W M PHILLIPS *Trans Electrochem Soc* 59 (preprint) 2 pp (1931)—In the St. Louis district Cr plated articles appear to have 'tarnished'. Investigation shows that where no Ni or insufficient Ni is first plated over brass or Cu articles, this type of failure can occur

C G F

Electrolytic production of hydrogen and oxygen. I I LINDVALL *Idee Times* 78, 380-92 (1930)—The advantages and disadvantages of the filter-press-type and bell-type electrolytic cells are discussed and the Knowles plant at Busil, Italy, is described in detail

I: 1 S

Leclanché-type primary cells. Inst Eng Standards Assoc, *Specification No* 397, 26 pp (Sept., 1930)—Dry cell specifications

I: 1 S

Asphalt composition vs. hard rubber in battery jars. M SHANK *Elec Mfg* 5, No 1, 37, 76 (1930)—The relative merits of these 2 materials are discussed

I: 1 S

The effect of electrical discharges upon thiocyanic acid. W GLUUD and CLARK DICKMANN *Ber ges Kohlentech* 3, 420-8 (1931), cf *C A* 24, 5811—The effect of elec discharges upon the reaction $HCNS(aq) \rightarrow HCN(aq) + S(\text{rhombic})$ was studied both silent and glow discharges being used. With the silent discharge, the yields were too low and the energy consumption extremely high. With the glow discharge, however, and with addn of H gas, about 80% of the HCNS was decomposed, 95-100% of which was recovered as HCN. Varying quantities of SO_2 , H_2S or S are also obtained, depending upon the exptl conditions. Elec. efficiency is very low, the energy consumption being about 675 watt hrs per mol. of HCN as compared with a theoretical requirement of only 6.87 watt hrs

I: 1 S

Electrical gas purification and filtration. R SEILIGER et al *Trans 2nd World Power Conference* (Berlin) 1, 492-504 (1930)—A review

AIDEN H I MERRY

Mercury-arc power rectifier auxiliaries and accessories. I S WATERMAN *Gen Elec Rev* 34, 228-31 (1931)—The rectifier depends on the successful functioning of its auxiliary app. The exhaust system is important, as a high degree of vacuum must be maintained. It consists of a vacuum valve, Hg condensation pump, receiver tank, motor-driven rotary vacuum pump and the vacuum measuring and regulating devices. The contact making gage is made of three parts: (1) thermal vacuum detector, (2) vacuum regulator and (3) source of excitation. The vacuum regulator is composed of the following parts all mounted in one case: (1) indicating motor, (2) regulator contacts and contact hammer, (3) small a c motor. The regulator and its auxiliary relays perform the following functions: (1) start and stop the rotary pump at desired pressures, (2) remove the rectifier from service if the pressure becomes excessive, when the unit is equipped with automatic control and (3) sound bell alarm if pressure be-

body such as Cu having a coating of a compd. such as Cu_2O placed in a suitable container with Hg under pressure. A layer of Pb foil also is used.

Electrodeposition of gold and silver. ASSOCIATION OF ELECTROLYTIC DEPOSITORS IN VEROVA, CHIEF. Ger 519,268, Dec 12, 1928. The electrolyte is AuCl_3 or AgCl dissolved in a salt soln. e. g. in an alkali or alk. earth hydroxide soln. The method may be applied in refining the metals or in electroplating. Suitable electrolytes are AuCl_3 and KCl 140 g., or AgCl 40 and KCl 320 g., in 1 l. of water, and suitable conditions are a voltage of 0.1-0.3 and a c. d. of 100 amps. per sq. m.

Chromium plating. JOHN L. K. McCLELLON and BENJAMIN W. GUERHART (to General Electric Co.) U. S. 1,797,367, March 21. A Cr plating bath is prep'd contg. $\text{Na}_2\text{chromate}$, Cr chromate, HCl and water. Cl. C. 4, 24, 302.

Electroplating iron and steel with other metals such as nickel and chromium. METALS PROTECTION CORP. Fr 693,066 May 9, 1930. See 1,774,219 (C. 1, 24, 4003).

Producing coherent metal layers on crystals. HANCOCK CHEMICAL AND MANUFACTURING CO. Ger 519,110, Dec 18, 1927. The crystals are subjected simultaneously to superficial reduction and electrodeposition, with the result that an alloy layer is formed on the cryst. surface. Thus, a layer of cryst. Cu_2O , for use in a dry rectifier, may be treated with an acidified soln. of $\text{NiSO}_4 \cdot (\text{NH}_4)_2\text{SO}_4$ undergoing electrolysis.

Covering metals. LEONARD J. KRAMER and VINCENZO RIZZI (to General Electric Co.) U. S. 1,797,367, March 21. In coating metals with compds. or mixts. known under the designation of W. Irons and Mo. bromides, the iron is obtained by electrolysis of solns. of salts of meta- or polytungstic acids, or salts of meta- or poly molybdic acids on the metal dipped into the soln. and used as cathode. The use of $\text{BaW}_2\text{O}_7 \cdot 9\text{H}_2\text{O}$ and $\text{BaW}_2\text{O}_7 \cdot 9\text{H}_2\text{O}$ is described.

Depolarizer of an app. especially applicable for the decomposition of NaCl . JAMES N. SMITH (to Westinghouse Electric & Mfg. Co.) U. S. 1,797,367, March 21. Structural electrode.

Electrolytic cell suitable for electrolysis of water. JAMES N. SMITH (to Westinghouse Electric & Mfg. Co.) U. S. 1,797,367, March 21. Structural features.

Method and apparatus for sterilizing water in thin layers by electrolysis. PATRICK M. R. SALLIS. Ger 519,418 July 17, 1926.

Metallized asbestos diaphragm for electrolytic cells. KARL WOLINKE. Ger 468,452, Mar 17, 1927.

Metals from ores. DUMREY RESEARCH CO. Fr 696,418, April 2, 1930. A process for the elec. heating and fusing of conducting materials such as metals reduced from their ores is described. The material is introduced from below into a column which opens on to the hearth of a furnace, the top of the column being used as electrode to establish an arc between it and the electrode of the furnace. The temp. is regulated by varying the speed of feeding.

Aluminum. FRIEDRICH HARTMANN. Fr 695,197, May 15, 1930. Al is obtained by reducing Al compds. such as bauxite in an elec. furnace in the presence of W, whereby an alloy of Al and W is obtained, from which the W is removed by a swarting or liquidation operation or an electrolytic refining below the m. p. of Al.

Magnesium. ALBERT C. JESSUP. Fr 696,218, May 28, 1930. Mg is obtained by making an incompletely dehydrated mixt. of KCl and MgCl_2 in the proportion of 2 or 3 mols. KCl to 1 mol. MgCl_2 , electrolyzing the mixt. until the mixt. contains at least 3 mols. of KCl for one of MgCl_2 and submitting the mixt. thus practically dehydrated to a second electrolysis. The no. of volts for the 2nd electrolysis is preferably greater than for the first.

Ammonium compounds. PAUL BAERWALD and HENRYK GOLDMANN. Ger. 521,238 July 17, 1927. In the manf. of NH_3 compds. by subjecting a mixt. of N, H and other gases to elec. discharge, the tension is alternately raised and lowered during the process. Thus, a mixt. of N, H and HCl may be subjected to an electrostatic field, the tension of which is varied every few secs. between 100,000 and 120,000 v.

Dioxides of titanium, zirconium, hafnium and thorium. IMPERIAL CHEMICAL INDUSTRIES, LTD and J. L. CLARK. Brit 369,618, Sept 7, 1929. An alk. soln. or suspension of raw material contg. a compd. of the metal, such as a suspension of ilmenite, is electrolyzed to deposit impurities such as Fe in the metallic state (suitably under 20 atm. pressure of H or air at 90-120°) with an insol. anode of C, Ni or Cr steel. Various details and examples are given.

Oxidized layers on aluminum. RICHARD F. MILLER. *Fr. 673,107*, May 24, 1930. In the formation of oxidized layers on Al, oxides of other metals such as Fe are incorporated during the operation. The result may be obtained by electrolysis, the electrolyte containing salts of Fe in the colloidal state, or the Al may be alloyed with metals to form the oxides such as Fe, Si and Cu.

Enamelled designs on metal surfaces. HUGO DEBACH. *Ger. 516,607*, Aug. 27, 1927. The metal surface is prepped chemically or electrochemically with layers which give the enamel the coloring of metal compounds. These layers are then covered with a transparent enamel or glaze. The whole is then heated until the latter layer fuses. Several examples are given. In one a Cu surface is galvanically silvered and then galvanically coppered, thickly for red parts and thinly for green parts. Electrolysis in a bath of 25% NaNO₃ and 1% Na₂CO₃ leaves the plate with a deposit of Cu, which is shaded off into the desired light shade and color effects by a rotating or scraping tool. The coloring effect may be enhanced by under-glaze pigments. The plate is then covered with glaze powder and heated in a muffle furnace.

Hydrogen peroxide. I. G. FARRINGTON A-G. *Fr. 606,150*, May 27, 1930. An app. is described for the production of H₂O₂ by electrolysis under pressure. Hg electrodes comprising several trays filled with Hg or amalgam and placed one above the other in a common pressure vessel are used.

Hydrogen peroxide. I. G. FARRINGTON A-G. *Fr. 606,273*, June 2, 1930. H₂O₂ is produced by the cathodic reduction of O, the O being controlled continuously or periodically or the H present being eliminated continuously or from time to time, preferably by catalytic combustion, or both methods are combined. An automatic analysis app. mounted in the current of O is used.

Hydrogen peroxide. I. G. FARRINGTON A-G. *Fr. 606,285*, June 3, 1930. See *Ger. 514,172* (C. A. 25, 1167).

Hydrogen peroxide. KALL-CHENET A-G. *Fr. 606,371*, June 1, 1930. H₂O₂ is submitted to an electroosmotic treatment for purification and concentration.

Bleaching shellac. OLGA MYLO NIE ROSENBERG. *Ger. 512,762*, April 20, 1930. Addn. to 517,096 (C. A. 25, 240). A soln. of crude shellac is mixed with a soln. of a fluoride, e. g., NaF or NH₄F, and the mixt. is electrolyzed.

Electric furnace. JEWELL I. BRANN. *Ger. 512,762*, Mar. 20, 1930. Addn. to 507,074 (C. A. 25, 473).

Electric furnace. W. C. HERBERT G. M. E. (Ernst Haagen and Edvard Killus, inventors). *Ger. 517,401*, Mar. 31, 1930. Means of regulation are described.

Electric furnaces. SHIMKIN-PLANTZ. *Fr. 606,116*, May 27, 1930. Electrodes for electric furnaces are made by packing powder or granular mineral coal, which has not been or has only been partly carbonized, into a vessel and heating it in the vessel until completely carbonized. Pressure may be applied to the coal if much gas is evolved during the heating. A flux such as asphalt, bitumen or resin may be added to the coal. The electrodes may be made in the furnace during its use. *Cl. C. A. 25, 1744.*

Electric resistance furnace. ART-GES. BROWN, BOYER & CO. *Swiss 143,509*, Nov. 29, 1929.

Induction furnace. PETER KAMPS and HERMANN ZIEFERTICK. *Ger. 516,539*, Mar. 10, 1930.

Electrode smoking furnace. ART-GES. BROWN, BOYER & CO. *Ger. 521,208*, July 12, 1929.

Rotary electric furnace suitable for reducing iron ore. B. M. S. KALLING and C. VON DELWIG. *Brit. 347,117*, Dec. 10, 1928. Structural and elec. features.

Electrolytic furnace. DOW CHEMICAL CO. *Brit. 339,833*, Jan. 11, 1930.

Electric annealing furnace with protective gas filling. ALGERMINE ELECTRICITÄTS-GES. *Ger. 519,231*, Feb. 16, 1927.

Electrically heated apparatus for concentrated sugar juices, cracking oils or other purposes. R. A. CARLTON. *Brit. 339,649*, Sept. 10, 1929. Various structural features are described of an app. comprising a plurality of closely spaced tubes connected at their ends by bends to form a continuous conduit through which the fluid to be heated is circulated and which is electrically heated.

Electrically heated apparatus for determining the limiting creep stress of materials at selected temperatures. D. COLVILLE & SONS, LTD., W. BAKER and W. E. BARDGETT. *Brit. 339,840*, Feb. 26, 1930. Structural features.

Electric heating unit suitable for immersion in water. CLARENCE M. BRADNAX and JOSEPH LONNO. *U. S. 1,797,712*, March 24, 1930. Structural features.

Electric air heater. WALTER W. BROWN and HLOYD B. STONE U. S. 1,797,519 March 24

Signal device for electric smelting furnaces. ALLGEMEINE ELEKTRIZITÄTS GES. Ger. 521,290, July 20, 1929. Molten metal leaking from a faulty melting pot or the like operates an alarm.

Carbon electrodes for electric furnaces. ROBERT SUCHY, KARL STAIB and WILHELM MOSCHEL (to I. G. Farbenindustrie A. G.) U. S. 1,797,800, March 24. See Fr. 665,620 (C. A. 24, 794).

Electrode of very wide section for electric furnaces. SOC. ELECTRO-METALLURGIQUE DE SAINT-BÉRON. Fr. 696,319, May 30, 1930.

Electrode for electrolytic apparatus such as cells for decomposing water. JAMES N. SMITH (to Westinghouse Elec. & Mfg. Co.) U. S. 1,797,375, March 24. Structural features.

Electrodes for electrolyzers. GEORGE F. JALBERT. Fr. 696,788, Sept. 16, 1929. The electrodes are coated with a mineral or org. colloidal substance, conductive or not, which prevents direct contact of the electrolyte with the metal of the electrode, whereby the polarization is diminished.

Soluble anodes for use in electrolysis. SIEMENS & HALSKE A. G. (Victor Engelhardt, inventor). Ger. 521,009, May 17, 1928. The anodes are made in the form of very thin long or endless bands which are kept in constant motion in the electrolyte.

Electrodes for electric purification of gases. SIEMENS-SCHÜCKERTWERKE A. G. Fr. 695,777, May 17, 1930.

Electrodes for secondary elements. I. G. FARBENIND. A. G. (Karl Ackermann, inventor). Ger. 519,456, Dec. 19, 1929. Addn. to 491,498 (C. A. 24, 2066). The porous metallic material used in the method of Ger. 491,498 is made of Cu-Ni alloy, e. g., an alloy containing Cu 50 and Ni 50%. Cf. C. A. 24, 4469.

Depolarization electrodes. JOHANNES M. SCHMIDT and DIAMCO A. G. Ger. 510,148, Apr. 5, 1930. In the manufacture of compressed depolarization electrodes, the active mass is mixed with org. fibers, e. g., ramie, in order to prevent swelling of the electrodes when in use.

Device for joining carbon furnace electrodes. SIEMENS-PLANAWERKE A. G. FÜR KÖHLEFABRIKATE (Josef Soppa, inventor). Ger. 520,394, Nov. 21, 1928.

Spark-discharge apparatus for generating a bleaching gas for flour, etc. FRITZ STUTZ. Ger. 521,360, Dec. 23, 1929. Addn. to 510,419 (C. A. 25, 1008).

Electric apparatus for testing gases. EDUARD SENFFER. Ger. 516,719, June 6, 1928. The app. is suitable for testing samples of crucible gas etc., from tubes of various pressures.

Apparatus for purifying gases. SIEMENS-SCHÜCKERTWERKE A. G. Fr. 696,117, May 20, 1930. The pipes for the sweeping gases for the insulators in elec. gas purification app. are disposed in the hot gaseous current passing through the app. whereby the sweeping gases are heated.

Gas purification. METALLGES. A. G. Fr. 696,226, May 30, 1930. Preliminary or premature ionization of gases to be purified by elec. means is suppressed by means of a screen disposed transversely in the current of the crude gases, the face of the screen being disposed above the source of the gas in the immediate neighborhood of the flame. The screen may be provided with cooling means.

Removing chemically active gases from oil containers, in electric apparatus, etc. SIEMENS-SCHÜCKERTWERKE A. G. Ger. 517,842, Aug. 7, 1928. The app. contains chambers with gas-absorbing solvents.

Electric switches and fuses. J. LANN. Brit. 339,748, Nov. 7, 1929. A tube filled with Ne, He or other rare gas, and constituting a glow-discharge tube is arranged across the terminals of a switch or fuse so that the tube glows when the switch is open or when the fuse blows, indicating the location of the switch or fuse in the dark.

Loading submarine signalling cables. STANDARD TELEPHONES & CABLES, LTD., and J. R. VESEY. Brit. 339,980, Sept. 18, 1929. Mech. features.

High-power incandescent lamps, tungsten arc lamps, x-ray tubes, rectifiers, etc. PATENT-TREHAND-GES. FÜR ELEKTRISCHE GLÜHLAMPEN (to General Elec. Co., Ltd.). Brit. 339,927, May 18, 1929. Structural features.

"Getter" for tungsten incandescent lamps. EGYENLETI IZZÓLÁMPÁ ÉS VILLAMOS-ÁRÁK RÉSZÉNYTÁRSASÁG. Brit. 339,798, Dec. 18, 1928. As a "getter" there is used an inorg. water-absorbing substance gaseous at ordinary room temp. such as SiF₄, AsF₃, or a fluoride of P. The getter may be introduced as such or may be formed in the bulb, as by decomposition of K silicofluoride.

1930 Bleach out layers, sensitized with thiourea derivs., are fixed by destroying the sensitizer with SO_2 and, if necessary, removing the decompos. product from the layer.

Photographic bleach-out process. I G. LARSEN AND A G. FR 695,426, May 10, 1930 The binding agent for the colors in the bleach-out process is composed of a mixt. of substances which dissolve the colors with substances which do not dissolve the colors and which swell in the fixer to be used in each case. Cellulose ethers or esters may be used as the substances capable of swelling in the fixers (CaCl_2 , other aromatic hydrocarbons benzene or ethers) and also of high b p or cellulose esters such as nitrocellulose as the solvents for the colors.

Photographic reproductions. RUDOLF FISCHER FR 695,921, May 10, 1930 In the photographic reproduction of pamphlets, notes, etc., the actinic rays are prevented from traversing the emulsions support by coloring the latter itself or a special layer applied thereto with substances which optically are as clear as possible but absorb the actinic rays.

Photographic images. I G. FARBER AND A G. FR 696,641, June 4, 1930 B. chrome or polychrome photographic images are produced by one of the methods used in color photography, using esters or salts of esters of leuco compds. of dyes as substances sensitive to light. The sensitiveness to light may be increased by the addn. of substances such as colloids having an oxidizing action to light.

Photographic films. KALLE & CO A G FR 696,100, May 26, 1930 Photographic films, etc., particularly cellulose films sensitized by diazo compds., are protected against damp by coating them before use with a layer of varnish such as nitro- or acetyl cellulose varnish.

Photographic films. WLADIMIR REYOS FR 696,419, May 12, 1930 Sensitized layers on films are protected by a coating composed of a soln., in an appropriate solvent, of the substance composing the support or of a like substance capable of uniting intimately with the support, or such a substance in a liquid state. Thus a soln. of celluloid in AcOAm may be used on celluloid films.

Metal cinematograph film. FREDERICK RAD6 GER 517,517, Nov 15 1927 The film consists of rolled metal, e. g., Al, with a coating of cellulose, viscose, etc., which is given the sensitized coating in the usual way.

Two-color cinematographic films. J. E. THORNTON BRT 339,977, Aug 10, 1929 In a modification of the process of film manuf. described in Bnt. 316,338 (C. A. 24, 1548), feed perforations are made in the blank film before it is fed to the printing machine. Various details of manuf. are described.

Coloring films. MULTICOLOR FILMS INC BRT 339,971, Sept. 17, 1929 See FR 684,354 (C. A. 24, 3248).

Sound films. I G. FARBER AND A G. FR 695,831, May 19, 1930 In making sound films in colors, the sound band is first produced in known manner by printing, developing and fixing, and afterward the colored image is produced in known manner. The sound band is protected against modifications by an impermeable layer.

Patterns for printing paper, etc. CONRAD LANG AND ARTHUR LANG SWITZ 133,800, May 4, 1929 The pattern template is placed on a light sensitive layer which is protected from the action of the printed colors, prior to the copying and developed.

Photographic prints in colors. ANITA TRUST BRT 340,044, Oct. 1, 1928 In producing prints in natural colors with use of 3 sensitive emulsions each dyed with a transparent dye, which are printed from color record negatives, developed, hardened and the unhardened gelatin and silver image removed, auxiliary dyes which are easily removed after exposure are added to the quicker printing emulsion layers to equalize the times of exposure of the layers.

Apparatus for developing blue prints. BRUNO SPANGLER G M F H GER 517,668, May 1, 1929

6-INORGANIC CHEMISTRY

A R MIDDLETON

Surface chemistry of hydrates. I V R DAMERELL J Phys Chem 35, 1061-7 (1931)—Hydrous Al oxide forms the trihydrate on standing in H_2O contg. small quantities of NH_3 . Over drying agents the hydrate particles lose chemically bound H_2O in proportion to the surface. The difference in H_2O content of ground and unground hydrargillite is due to increase in surface. S LEXNER

Hydrates of the alkaline earth oxides C NOGAREDA *Anales soc. esp. fis.*

quim 29, 33-64(1931)—Systematic dehydration of 2 lime hydrates of different origins, containing a large excess of free moisture, was carried out. From the dehydration diagram obtained there was no evidence of the presence of hydrates of $\text{Ca}(\text{OH})_2$, but the mol vol values of the free H_2O pointed to the existence of a labile $\text{Ca}(\text{OH})_2 \cdot \text{H}_2\text{O}$. By the usual methods $\text{Ca}(\text{OH})_2$ was obtained in macrocrystals, but without any hydrates. $\text{Sr}(\text{OH})_2 \cdot \text{H}_2\text{O}$ and $\text{Ba}(\text{OH})_2 \cdot \text{H}_2\text{O}$ were obtained by allowing their hot, satd solns to stand. Dehydration diagrams and detn of mol vol of H_2O of crystn showed clearly the presence of a little stable $\text{Sr}(\text{OH})_2 \cdot 8\text{H}_2\text{O}$ and $\text{Ba}(\text{OH})_2 \cdot 8\text{H}_2\text{O}$ and a very stable $\text{Sr}(\text{OH})_2 \cdot \text{H}_2\text{O}$ and $\text{Ba}(\text{OH})_2 \cdot \text{H}_2\text{O}$, showing again the great similarity of Sr and Ba compds.

E. M. SYMMES

Hydrates of the alkaline earth peroxides. II. Their constitution. C. NOGAREDA. *Anales soc. españ. fis. quim.* 29, 131-45(1931)—The d and mol vol of highly purified CaO_2 , SrO_2 , and BaO_2 octahydrates were detd. From previous work (*C. A.* 24, 3721) on d and mol vol of hydroxides of these metals the av mol vol for $8\text{H}_2\text{O}$ could be calcd, and by deducting this value from the values for the above peroxide hydrates the mol vol of the peroxides was obtained approx. The values found were: $\text{CaO}_2 \cdot 8\text{H}_2\text{O}$ d 1.700, M_r 127, $\text{CaO}_2 \cdot M_r$ 21.2, $\text{SrO}_2 \cdot 8\text{H}_2\text{O}$ d 1.951, M_r 135.1, $\text{SrO}_2 \cdot M_r$ 28.3, $\text{BaO}_2 \cdot 8\text{H}_2\text{O}$ d 2.201, M_r 136.8, $\text{BaO}_2 \cdot M_r$ 30.0. Comparison with the corresponding mol vol of the monoxides gives for the second O atom a vol about 0.25 normal, s_e , 2.9 to 3.0 (Biltz, *C. A.* 25, 628). All peroxide compds show clearly the Resenfeld reaction. Extn tests with Et_2O by the Willstätter method to detect H_2O_2 were neg. The isotherm of breakdown of the hydrate can be followed up to a loss of 7.5 H_2O , and then decompn starts. From the course of the dehydration curve no conclusions can be drawn as to the existence of other hydrates. All properties of the hydrates tested point to the

structure $\text{X} \begin{array}{c} \diagup \text{O} \diagdown \\ \diagdown \text{O} \diagup \end{array} 8\text{H}_2\text{O}$. In the breakdown of the hydrates the 0.5 H_2O remaining is

driven off only above 200° ; this can be explained by the assumption of a partial hydrolysis of the peroxide hydrate with formation of a hydroxide. The ratio $\text{XO}_2/\text{X}(\text{OH})_2$ in the dehydrated product was about 5:1. These hydroxides may act to stabilize the peroxides and explain the existence of CaO_2 obtained otherwise only under high pressure.

E. M. SYMMES

Iodates of titanium. PRIVADA R. RAY AND HARIBOLA SAHA. *Proc. 15th Indian Sci. Congr.* 1928, 169-70—By the action of HIO_3 dissolved in HNO_3 (1:3 by vol) upon freshly prepd. titanous acid dissolved in the same solvent, a basic iodate of Ti has been obtained. The individuality of the substance in question has been established by the fact that the same product was obtained from the liquid phase of varying concn. It is preferable to regard the compd as a complex tetrahydroxyiodatotitanic acid for several reasons. It is converted by heating to 100° to the proper basic Ti iodate. Salts of complex hydroxyiodatotitanic acid have also been prepared.

E. J. C.

A contribution to the study of the formation of salts with glycine. J. V. DUNSKI AND A. RABAS. *Spisy vyšševé přírodovědeckou fakultou Masarykovy Univ.* No. 123, 3-18(1930), cf. *C. A.* 24, 4722—Glycine (I) 7.5 g in 100 cc H_2O was treated with 4.07 g ZnO at $20-40^\circ$, filtered and crystd., it yielded 2.8 g $(\text{NH}_4\text{CH}_2\text{COO})_2\text{Zn} \cdot \text{H}_2\text{O}$ (II). II is a white, cryst. nonhygroscopic substance, sol in H_2O 1:10, turns yellow at $240-260^\circ$, decomposes at 310° , hydrolyzes at $65-70^\circ$ in H_2O with the formation of $\text{Zn}(\text{OH})_2$. II was also prepd. by treating I with ZnCO_3 or $\text{Zn}(\text{OH})_2$; II could not be prepd. by treating $\text{NH}_4\text{CH}_2\text{CO}_2\text{Ag}$ with ZnCl_2 . The mother liquor from II was further evapd. at 50° , yielding 2.9 g of a white cryst. mass $(\text{NH}_4\text{CH}_2\text{COO})_2\text{Zn} \cdot 3\text{ glycine} \cdot \text{H}_2\text{O}$ (III). III melts with decompn at 245° , is sol in H_2O 1:6 and nonhygroscopic, the solns are neutral. Aq. solns of III crystallize, forming first I and then III in the last fractions. III was also prepd. by treating 4.07 g ZnO in 100 cc H_2O with glycine (15 g) and evapg. untill crystals began to appear. $\text{ZnCl}_2 \cdot 3\text{ glycine}$ (IV) was prepd. by dissolving 3 g glycine in 20 cc H_2O , adding 2.73 g anhyd. ZnCl_2 in 5 cc H_2O and permitting free crystn to occur for 10 days. The yield was 1.9 g, m (with decompn) 228° , hygroscopic, acid to litmus and having a bitter taste. $\text{ZnCl}_2 \cdot 2\text{ glycine} \cdot 2\text{H}_2\text{O}$ was prepd. by treating 6.81 g anhyd. ZnCl_2 in 3 cc H_2O with 7.5 g glycine in 18 cc hot H_2O . The filtered soln. immediately gave glassy crystals in 100° and decompd. at 230° . $\text{ZnBr}_2 \cdot 2\text{ glycine} \cdot 2\text{H}_2\text{O}$ was prepd. by mixing 2.25 g ZnBr_2 and 2.25 g glycine in 7 cc H_2O . After standing, crystals (transparent) appeared, they m 98° . $\text{ZnSO}_4 \cdot \text{glycine} \cdot 5\text{H}_2\text{O}$ (II 6 g) was prepd. from 11.5 g $\text{ZnSO}_4 \cdot 7\text{H}_2\text{O}$ and 3 g glycine in 10 cc H_2O by allowing crystn to proceed for 14 days. The white transparent crystals lose 4 H_2O at $50-60^\circ$ and leave $\text{ZnSO}_4 \cdot \text{glycine}$ after drying at $80-125^\circ$. The crystals are stable in air, nonhygroscopic (m 82°) and decompose at 320° . (NH_4 -

The formation of green manganese sulfide in presence of NH_4HS , NH_4OH and NH_4S on the transition of the red sulfide to the green. G. ROSE, *Landmann*, 7, *Angew. Chem.* (Berl.) 193, 276 (1920) cf. C. A. 18, 598.—The effect of NH_4HS , NH_4OH and $(\text{NH}_4)_2\text{S}$ on the transformation of the red to the green MnS was studied by observations over 6 1/2 years. The presence of a slight excess of NH_4OH , up to 1 m. l. per atom of Mn , greatly decreases the time required for the initial appearance of the green form. A larger ratio of NH_4OH has the opposite effect. Changing the ratio of NH_4HS or of $(\text{NH}_4)_2\text{S}$ used for pptn. does not affect the initial appearance of the green. An excess of NH_4OH also does not increase the total rate of transformation. Independently of the green transition and simultaneously with it, but somewhat faster, occurs a transition to the orange sulfide. This modification, like the red, is amorphous while the green is crystalline. The presence of NH_4OH also hastens the red to orange change. The interpretation of these results shows that the initial reaction in the pptn. reaction yields the complex $\text{Mn}(\text{OH})\text{HS}$ from which the stable green MnS is formed. In the presence of a large excess of NH_4OH , the equl., $\text{Mn}(\text{OH})_2 + 2 \text{NH}_4\text{HS} \rightleftharpoons \text{Mn}(\text{HS})_2 + 2 \text{NH}_4\text{OH}$, is shifted to the left. Green MnS is formed from $\text{Mn}(\text{HS})_2$ only by evolution of H_2S , which accounts for the slow transformation.

H. P. JOHNSON

Lead oxides and their reducibility with carbon. P. VA RUDR, J. Fuss, *Proc. Chem. Soc.* 62, 1419-42 (1930).—The existence of 2 classes of lead oxides is established. To the first class belong oxides forming a series of solid solns. α , from PbO to $\text{PbO}_{0.5}$ (content of active O from 0 to 4.47%). The color of these oxides varies from light to dark brown. To the second class belong lead oxides forming a series of solid solns. β , from $\text{PbO}_{0.5}$ to $\text{PbO}_{0.25}$ (active O content from 2.2 to 6.24%), and solid solns. γ , from PbO to $\text{PbO}_{0.25}$ (active O content from 0 to 0.17%). Between the β and γ series there is an interval from $\text{Pb}_{0.5}\text{O}$ to $\text{PbO}_{0.25}$. The color of the oxides varies from dark brown in $\text{PbO}_{0.5}$ to red in $\text{PbO}_{0.25}$ and from light brown in $\text{PbO}_{0.5}$ to yellow in PbO . The character of the bond of active O in all these oxides is of the zeolite type. Reduction with solid C is highly exothermic in the case of oxides rich in active O. Reduction takes place in 2 stages of which the first probably corresponds to formation of a suboxide and the other to complete reduction. Oxides rich in active O contain some H₂O in the form of solid solns. The change of red lead oxide to yellow is not accompanied on the heating curve by any thermal effect.

S. L. MADORSKY

Action of aqueous sodium polysulfide solutions on metallic copper. K. FISCHBECK AND E. ELSEN, *Z. Elektrochem.* 34, 642-5 (1928).—The times taken by sulfide solns. of various concns. to form Cu_2S layers on metallic Cu of a standard color were measured. With solns. contg. virtually Na_2S_2 the velocity of thickening of the layer is proportional to the polysulfide content of the soln. and inversely proportional to the thickness of the layer. An analogous result is obtained with S vapor. The method of observation previously used (Fischbeck, C. A. 20, 3152) cannot be applied to the interaction of S and Cu. Some observations are made on sulfide solns.; the pentasulfide ion is probably more highly colored than the tetrasulfide ion.

B. C. A.

The action of hydrogen sulfide on potassium chromate solution. H. B. DUNN, CLIFF AND CHAMAN, L. SONT, *Proc. 15th Indian Sci. Congr.* 1928, 167.—In dil. solns. (2.15% K_2CrO_4) the color changes from yellow to green with evolution of heat. The

Cr is recovered quantitatively as hydroxide and the K is obtained as K_2S_2 and K_2SiO_4 .
E J C

The interaction of sulfur dioxide and ammonia. G SAMBAMURTI AND M N L NARASIMHAM *Proc 15th Indian Sci Congr* 1928, 169—Dry SO_2 and dry NH_3 interact to form 2 isomers of the constitutions $HSNS(O)ONH_4$ and $HSNS(O)_2NH_4$. No details are given.
L J C

Reactions involving hydrogen peroxide, iodine and iodate ion. III The reduction of iodate ion by hydrogen peroxide. HERMAN A LIEBHAFSKY *J Am Chem Soc* 53, 896-911(1931) cf *C A* 25, 1430—The rate of reduction of iodate ion by H_2O_2 is expressed within $\pm 5\%$ by $-d[IO_3^-]/dt = 2.6 \times 10^{-4} [H_2O_2][IO_3^-] + 129 \times 10^{-4} [H^+][IO_3^-][H_2O_2]$. The first step in the reduction of iodate may occur in the following ways: $IO_3^- + H_2O_2 \rightarrow IO_2^- + H_2O + O_2$, $IO_3^- + H_2O_2 + H^+ \rightarrow HIO_2 + H_2O + O_2$ followed by $H^+ + IO_2^- \rightarrow HIO_2$, $HIO_2 + H^+ + I^- \rightarrow 2HIO$, $2HIO + 2H_2O_2 \rightarrow 2I^- + 2H^+ + 2H_2O + O_2$, $HIO + I^- + H^+ \rightarrow HIO + I_2$. At the beginning of every expt there is an interval of time during which the rate of formation of I_2 differs from that required by the equation $dI_2/dt = k_1[H_2O_2][IO_3^-] + k_2[H^+][H_2O_2][IO_3^-]$. This time interval is necessary to establish the concns of intermediate products required for the reduction of the iodate ion. DON BROUSE

The influence of coordinately bound groups on the properties of the central iron atom in iron cyanide compounds. OSKAR BAUMSCH *Biochem Z* 232, 35 (1931)—Reagents are described for the identification of complex Fe ions. Benzidine and guaiacols are oxidized by ferric as well as autoxidized ferro-aquo salts, but 5 aminouracil reacts only with the latter at pH 3.8-8.0. Other compds of the same type as 5 aminouracil give the same reaction and serve as very sensitive reagents for the ferro-aquo salts. From fermentation isobutylalcohol a N contg substance was obtained which also reacts with freshly prepd ferro-aquo salt solns. Ferric ions can be demonstrated in the presence of complex ferric ions by means of thioglycolic acid, while nitrosophenol distinguishes ferrous ions. For the identification of ferro-aquo ions the reaction with dithiooxamide has been employed. The ferro-aquo salt soln is oxidized by 0.01 N H_2O_2 directly only in an acid medium. The ability of a series of substances to reduce ferro-aquo salt solns and to inhibit the reaction with dithiooxamide is recorded in a table.
S MORSELLIS

The influence of pH upon the formation and decomposition of the chloro derivatives of ammonia. ROBERT M CHAPIN *J Am Chem Soc* 53, 912-20(1931), cf *C A* 23, 4158—This paper is a study of factors influencing the kind and amt of products obtained by the chlorination of NH_3 . Cl_2 attacks NH_3 ion only after conversion into $HClO$. H ion induces the formation of NH_4^+ ion from NH_3Cl and NH_4Cl , particularly below a characteristic pH , the $HClO$ resulting ($NH_4Cl + H^+ + H_2O \rightarrow NH_4^+ + HClO$) reacts to produce a more highly chlorinated deriv. OH^- ion induces the formation of Cl^- ion from NH_3 or NH_4Cl , particularly above a characteristic pH , with N_2 and ClO^- ion as the principal assocd products. With increasing diln N_2O , NO_2^- ion and NO_3^- ion also appear. C is unable to confirm the statement that some O_2 is formed by the passage of Cl_2 into NH_3 water. Cf *Chem News* 5, 246 (1862).
W C FERNELIUS

The system water-potassium nitrate-calcium nitrate at 25° . MOHAMMAD A HAMID AND RAM DAS *Proc 15th Indian Sci Congr* 1928, 169—On account of the relatively great electroaffinity of the NO_3^- ion, the formation of double nitrates is found to occur only when one of the metals has a valency greater than two or when one metal is of high at wt and vol. The unexpected formation of a double nitrate of K and Ba, $K_2Ba(NO_3)_4$ (Wallbridge, *Am Chem J* 32, 251(1904)) led to a consideration of the possibility of the existence of similar double nitrates of K with other alk earth metals. In the case of Ca no evidence of double-nitrate formation has been found from an investigation of the system $H_2O-KNO_3-Ca(NO_3)_2$ at 25° . A no of hydrates of $Ca(NO_3)_2$ have been reported in the literature. H and D's observations that the only stable hydrate at 25° is the tetrahydrate, $Ca(NO_3)_2 \cdot 4H_2O$ agree with those of Basset and Taylor (*C A*, 6, 2044).
E J C

Pneumatolytic synthesis of silicates. I. C. J VAN NIEUWENBURG AND H B BLUMENDAL *Rec trav chim* 50, 129-38(1931)—On the basis of previous expts (*C A* 24, 5553) on the volatilization of SiO_2 by steam above its crit temp, silicates have been synthesized by placing the metal oxide and SiO_2 in sep crucibles and heating them in an autoclave with steam at temps of about 365° and pressures of 200 to 250 atm. In this way β -wollastonite, Sr metasilicate, alamosite, sillenite, tephrosite, antigorite and diopside have been prepd. Such syntheses may be very significant from

I. The Alkaline Earth Metals. By MAX SYBIL HURR. 316 pp. \$9.50 (Published in 1925) Vol. VI, Part I. Nitrogen. By EDWARD H. R. FRIDLAUX and HERBERT LAMBOURNE. 242 pp. \$10 (Published in 1928) Philadelphia: J. B. Lippincott Co.

7—ANALYTICAL CHEMISTRY

W. T. HALL

Recommended specifications for analytical reagent chemicals. W. D. COLLINS, GREGORY P. HAYTER II, V. FARR, J. V. FREEMAN, JOSEPH ROSIN, G. C. SPENCER AND EDWARD WICHMERS. *Ind. Eng. Chem., Anal. Ed.* 3, 221-4 (1931).—Specifications for reagents to be used in careful analytical work are given for 60% HClO_4 , $\text{CdCl}_2 \cdot 2\frac{1}{2}\text{H}_2\text{O}$, CrO_3 , $\text{Pb}(\text{NO}_3)_2$, MgO , HgCl_2 , Hg_2Cl_2 , Na_2NO_2 , $\text{c}(\text{CN})_6$, and $\text{UO}_2(\text{NO}_3)_2$. Corrections are given for published specifications for the correction to be applied in the gravimetric determination of SO_4^{--} for H_2PO_4 , $(\text{NH}_4)_2\text{HPO}_4$, BaCO_3 , MgCl_2 , $\text{K}_4\text{c}(\text{CN})_6$, $\text{K}_4\text{c}(\text{CN})_6$, $\text{K}_2\text{H}_2\text{PO}_4$, NaHCO_3 , Na_2HPO_4 , and $\text{Na}_2\text{S}_2\text{O}_8$. W. T. H.

A simple method for determining the concentration of definite colloidal solutions. OTTO LINSTEN and HERMANN BORCHERT. *Z. anal. Chem.* 83, 101-10 (1931), cl. C. A. 24, 1267.—To a solution containing small portions of liquid until the Tyndall or flocculent effect disappears. With the aid of standard curves it is then possible to test the content of the solution. Typical results are shown with mixtures of cholesterol and water and also eosin and water. W. T. H.

Applications of ceric sulfate in volumetric analysis. IX. Standardization of thiosulfate solutions. Determination of thiosulfate. N. HOWELL FURMAN and JOHN H. WALLACE, JR. *J. Am. Chem. Soc.* 53, 1281-8 (1931), cl. C. A. 24, 3720.— $\text{Na}_2\text{S}_2\text{O}_3$ solutions can be titrated iodometrically with standard $\text{Ce}(\text{SO}_4)_2$ solution, the starch iodide end point gives better results than the potentiometric method. Air must be excluded and if Fe^{+++} is present, it must be rendered inactive by adding PO_4^{---} . The initial volume should be 10 times that of the 0.1 N thiosulfate used and 0.15 g. of KI should be used for each 125 cc. of the initial solution. W. T. H.

Stabilizing hydriodic acid. HERMAN DUKOSKI. *Chemist-Analyst* 20, No. 3, 15 (1931).—The addition of 1 g. of red P to 1 l. of 11 N solution will stop the liberation of I₂ after just before using. W. T. H.

The use of polarographic methods in applied chemistry. J. HURROVSKY. *Chem. Listy* 24, 119-28, 417-8 (1930), cl. C. A. 24, 4200.—A résumé of the application of polarographic methods is given. FRANK MARSHALL

Some application of Wood's light. I. COLOMBIER. *Ann. chim.* 24, 89-96 (1931).—The fluorescence of various compounds when examined by ultra-violet light was examined with the following results: Eosin varies from dark to luminous over a pH range of 2.5-4.5. Erythrosin varies from dark to luminous over a pH range of 4-15. Fluorescein varies from dark to luminous over a range of 4-15. Quinine shows strong fluorescence at pH 5, decreasing sharply at first and then more slowly, becoming dark at pH 9. Acridine shows a sharp change from green to violet at pH 5. With umbelliferone fluorescence appears sharply at pH 0.5. With β -methylumbelliferone fluorescence appears sharply at pH 7. With β -naphthol fluorescence appears sharply at pH 9.5. None of these compounds is suitable for pH determinations, but the last 4 can be used as acidimetric indicators. Particularly with umbelliferone and β -methylumbelliferone, good results were obtained with strong acids and bases at dilutions up to 0.001 N and with AcOH and NH_3 at dilutions up to 0.01 N; also with carbonates by working at boiling temperature, or preferably by adding excess acid, boiling to expel CO_2 and titrating back with alkali. Acridine gives good results, but its use is advantageous only for NH_3 for which it can be used at dilutions up to 0.002 N. At higher dilutions, it is less sensitive than umbelliferone. β -Naphthol is of some interest, but its end point lies in the alk. range, and the use of umbelliferone is preferable. Determination of the acidity of wine by titration in ultra-violet light in the presence of umbelliferone is not satisfactory because of the strong absorption of ultra-violet rays by the wine; it can be carried out by microtitration (dilution of 1 cc. of wine to 250 cc. and titration with 0.001 N alkali), but the results are considerably lower than by titration in the presence of phenolphthalein because the end point of the latter is at a considerably higher pH (9-10 as compared with 6.5 for umbelliferone). The method gave excellent results for the direct titration of highly colored dye-house effluent with 0.1 or 0.01 N alkali. A no. of qualitative tests based on the formation of a fluorescent compound have been developed: (1) Detection of malic acid: add 1 cc. of 0.1% solution of resorcinol to a few cc. of solution or a few particles of powder to be tested, evaporate to dryness on a water

al. take up in 1 cc of 3 N H_2SO_4 (3 vols H_2SO_4 + 1 vol. water), heat 5 min. on the water bath with dil with 10 cc water and neutralize exactly with NaOH ; if malic acid is present there is a beautiful blue fluorescence in ultra-violet light, the test is sensitive to 1 mg. The reaction is not absolutely sp., as it is given by all β -ketonic acids; but the following are of rare occurrence and their other characteristics are quite different. (2) *Malic acid*—A few cc of the soln with dil KMnO_4 to a light yellow color. (3) *Malic acid*—Add 1 cc of 0.1% resorcinol and proceed as for malic acid. The reaction is sp. to 1 mg. The color of the fluorescence is different from that produced by malic acid and is readily distinguished therefrom, especially if a comparative test is made. (4) *Detection of resorcinol*—Add 1 cc of concd H_2SO_4 to a small amt of the sample, followed by 1 cc of acetoacetic ester. Sufficient heat is developed by the reaction. Observe directly. Resorcinol gives a beautiful violet fluorescence, which is intense even with 0.05 mg resorcinol. The reaction is sp., β -naphthol being the only other that giving a similar reaction, but the coloration is quite different and is produced by the addition of H_2SO_4 alone. (5) *Detection of β -naphthol*—Dissolve a few mg of the sample in concd H_2SO_4 and heat 2 min. on the water bath; in ultra-violet light there is a beautiful violet blue fluorescence if β -naphthol is present. The reaction is sp. and is particularly useful for distinguishing between α - and β -naphthol. (6) *Detection of orcinol*—In Crump's reaction (treatment with CHCl_3 and potash) after diln. there is produced a beautiful green fluorescence, which is much stronger in ultra violet light than in ordinary light and which readily makes possible detection of a fraction of a mg of orcinol.

A. PAPIEUX-COUTURE

Diphenylamine as an inside indicator for the determination of iron. CARROLL B. CORN. *Chemist-Analyst* 20, No. 3, 14 (1931).—Further commendation of this well-known method. Even with KMnO_4 titrations the addition of 4 drops of the indicator soln. has been found advantageous. W. T. H.

Solutions for colorimetric standards. IV. Some factors affecting the color of indicator solutions. M. G. MELLON AND G. W. FERNES. *J. Phys. Chem.* 35, 1023-43 (1931). Cf. C. A. 24, 2362.—Because of trouble in determining values with indicators from different sources the present work was undertaken to secure information on the effects of (1) the purity of the indicator, (2) the method of prep. stock solns of indicators and (3) aging, particularly with respect to the effect of light upon the color. Thirteen indicators were studied: methyl orange, tropaeolin OO, bromophenol blue, bromocresol purple, bromothymol blue, phenol red, bromocresol green, methyl red, cresol red, thymol blue, chlorophenol red, bromophenol red and cresolphthalein. As a result of a systematic spectrophotometric study, it is clear that the source of the compds., the way in which the solns are prep. and the effect of ultra violet rays are not without influence on the color of the solns. On account of the decolorizing effect of ultra-violet radiation on buffered solns containing indicators, the solns. to be used as standards should be protected from strong sources of photochemically active radiant energy. A uniform method should be followed in making up the color standard solns. Definite specifications should be established for indicators to be used as color standards. W. T. H.

The reduction of phosphomolybdic acid by monomethyl- p -aminophenol and its use in the quantitative determination of phosphorus. S. L. LEIBOFF. *J. Lab. Clin. Med.* 16, 495-9 (1931).—Monomethyl- p -aminophenol is a satisfactory reducing agent for phosphomolybdic acid in the determination of P. The reagent is stable in the presence of Na_2SO_4 and the color produced is stable within a wide range of acidity. E. R. MAIN.

Oxygen in aluminum and a method for its determination. TAKAYASU HARADA. *Anniversary Vol. Dedicated to Masumi Chikashige* 1930, 237-43.—Less than about 0.38% of O is present in solid soln in Al metal. If more O is present in a sample, it exists as Al_2O_3 at the boundary lines of the Al crystals. The method described depends upon heating Al shavings gradually to 750° in a strong stream of Cl_2 , which has been purified by passing the gas over hot C, it was proved that the CO formed had no effect upon the analysis. After 15 min. at 750° all of the Al should be removed as volatile AlCl_3 . The residue in the boat was washed with very dil. HCl and analyzed for Al_2O_3 after removing Fe and SiO_2 . W. T. H.

A qualitative color reaction for magnesium. J. V. DUBSKÝ AND ARN. OZÁČ. *Chem. Listy* 24, 492-3 (1930).— α , p -Dihydroxybenzeneazo- p -nitrobenzene gives a blue coloration with salts of Mg in an alk. medium. The color is not a compd. of Mg or a lake, but an adsorption of the dye upon the $\text{Mg}(\text{OH})_2$. The color is not sp.: Ni salts give an identical color, while Co salts give a different shade of the same color. The dye (0.5 g.) is dissolved in 100 cc. of a 1% alk. soln., the dye is very insol. in neutral water. One drop of the dye soln. is added to an unknown Mg soln. and 2 N NaOH is added. The alk. dye soln. is violet, upon adsorption by the $\text{Mg}(\text{OH})_2$ a distinct blue color

appears. It will detect 0.00002 mg. Mg. Thiazole yellow A in a 0.01% soln. is pale orange, upon absorption by $Mg(OH)_2$ it becomes rose and will detect 0.0002 mg. Mg per cc. Thiazyl orange R (S), benzopurpurin 4B (Gr-F), azo blue B and thiazine pure blue FF (C) give only vague color changes. FRANK MARSHALL

A method for the determination of thorium, uranium and potassium in specimens of stones and minerals. HENRIETTA GILSON. *Sitzb. Akad. Wiss. Wien. Abt. IIa*, 139, 181-93 (1930).—A modified C. T. R. Wilson tilted electroscope with ionization chamber attached was designed to use in detg. the U, Th and K contents of stones and minerals without even breaking the specimens. A careful study of the thickness correction was made by the use of artificial stones made of plaster of Paris with active substances mixed in. To det. U, Th and K it is necessary to make 3 measurements of the ionizing power, one without alcohol, one with glass or gypsum alcohol, (differentiated absorption of U and Th series radiations), and one with At. test for K. An indication of the accuracy of the method is the uniform decreasing activity with increasing basicity of the rocks. The method is probably accurate within 10% for U and Th but not so good for K. OLIVIER P. STUART

A sensitive test for molybdenum. JOSEF V. LACHYNA. *Chem. Zvesti*, 24, 165-6 (1930).—The test for Mo with K f t xanthate can be made more sensitive by using a deriv. of a higher alc. the cetyl alc. deriv. is the most sensitive. The soln. is prepd. by mixing a acid soln. of cetyl alc. in C_2H_5 with powd. KOH or a satd. alc. soln. of KOH. The unknown neutral soln. is treated with a few drops of cetyl xanthate and made acid with 1.0 N HCl. The presence of molybdic acid gives a dark violet color. In extreme diln. the color adheres to the suspension of cetyl alc. The limit of detection is 0.0005 mg. Mo per cc. and the limit of diln. 1:2,000,000. Mo may be detected in a drop of unknown liquid. Filter paper is impregnated with cetyl xanthate, dried, treated with a drop of unknown soln. and exposed to HCl fumes. The limit of detection is 0.0001 mg. Mo per 0.05 cc. The filter paper must be freshly prepd., as it keeps only about 24 hrs. The effect of inorg. salts upon the reaction is insignificant. W. gives no color reaction with xanthates, so Mo may be detected in the presence of W. Solns. with a high W content must be made strongly acid to insure a complete soln. of Mo and W. 0.01 mg. Mo in the presence of 0.5 g. W contained in 5 cc. water was detected by the reagent. The limit of diln. which still gave a color was 0.00002 mg. Mo per 0.05 cc. soln. FRANK MARSHALL

Lead in basic bismuth nitrate. N. SCHÖN. *Pharm. Weekblad*, 69, 277 (1931).—The H_2SO_4 test for Pb in basic Bi nitrate is capable of detecting 0.1%. A. W. DOX

Method for the determination of silver. R. STEVENS GIBBS. *Chemist Analyst*, 20, No. 3, 10-1 (1931).—Directions are given for detg. Ag as $AgCl$. The manipulation is somewhat different from that ordinarily recommended. W. T. H.

Modified gravimetric determination of manganese. FREDERICK G. GRANT. *Chemist Analyst*, 20, No. 3, 11 (1931).— $MnO(OH)_2$ can be pptd. in the presence of H_2SO_4 by adding $K_2S_2O_8$. Upon ignition Mn_2O_3 is formed. W. T. H.

Determination of zinc with pyridine and ammonium thiocyanate. C. FRANKLIN MILLER. *Chemist Analyst*, 20, No. 3, 8-9 (1931).—For the analysis of a Zn salt, take an aliquot contg. about 0.3 g. Zn, dil. to 500 cc. and add a soln. of 4 g. NH_4CNS and 2 g. pyridine in a little water. Filter, wash first with a soln. contg. 3 g. NH_4CNS and 5 cc. pyridine per l., next with a mixt. of 13 cc. 95% alc., 85 cc. water, 15 cc. pyridine and 0.1 g. NH_4CNS , then with 10 cc. of alc. mixed with 1 cc. of pyridine and finally with 15 cc. of ether contg. 2 drops of pyridine. Dry to const. wt. (10 min.) at 65-70°. The ppt. contains 19.25% Zn. W. T. H.

Colorimetric determination of minute amounts of cadmium. LAWRENCE T. FAIRBANK and LEON PRODAN. *J. Am. Chem. Soc.*, 53, 1321-3 (1931).—The method, applicable to the detn. of traces of Cd in org. material, depends upon pptn. by means of H_2S in the presence of 0.5 mg. Cu and 2 g. of Na citrate in a soln. which has been made $pH = 3$. The ppt. of CuS and CdS is dissolved in aqua regia, the soln. evapor. to dryness and again treated with H_2S under the same conditions as before and finally a third pptn. is effected at $pH = 2$ without any addition of citrate. This last ppt. is dissolved, the acid removed from the soln. by evapn. and the residue dissolved in a definite vol. of water. Aliquots are then dil'd., treated with KCN soln. and some H_2S water. The color of the ppt. is then compared with that of similarly-treated standard solns. under a flood of ultra-violet light. The color deepens on standing. W. T. H.

Colorimetric determination of tungsten. G. HAYN. *Z. anorg. Chem.*, 44, 217-8 (1931).—In the absence of NO_3^- , Fe^{+++} , Ti^{++++} , Ch , CrO_4^{--} , ReO_4^{--} and MoO_4^{--} , small quantities of W can be detd. satisfactorily by the reaction with hydroquinone in the presence of concd. H_2SO_4 . If Mo is present, the test fails because of its color

reaction with the reagent. In such cases *Rhodamine B* is a more satisfactory reagent. To be sure it also reacts with Mo salts but it requires 10-20 times as much Mo as of W to give the test so that it can be used when the quantity of Mo is smaller than that of W. Otherwise it is necessary to det. the Mo by a different method and apply a suitable correction. Directions are given in detail for carrying out these tests, which have already been described by others. W. T. H.

Estimation of copper in cyanide solutions. E. H. SMITH *Chemist-Analyst* 20, No. 3, 10(1931).—To det. Cu in a soln used for extg. Au from ores, take 100 cc. and destroy CN^- by adding 5 cc. of concd HNO_3 . Evap. to 20 cc., make acid with $AcOH$, add KI and titrate with $Na_2S_2O_3$. W. T. H.

Study of the electrolytic determination of copper in the presence of platinum ion. T. LEONARD KELLY and JOSEPH J. MULLOY *J. Am. Chem. Soc.* 53, 1337-41(1931).—Fales and also Popoff, have shown that sometimes Pt is dissolved off the anode and deposited upon the cathode during electrolysis and the question arises as to whether the deposited Pt will dissolve off the cathode when the Cu is dissolved in HNO_3 . In the 20 expts. here described, a little H_2PtCl_6 was added to the electrolyte after a considerable part of the Cu had been deposited. After the electrolysis, the gain in wt. indicated that practically all of the Cu and Pt had been deposited and, after dissolving off the Cu, the cathode weighed more than originally. In only 6 cases, however, was the final gain in wt. of the electrode equal to the total Pt in the soln. and in no case was there any evidence of any Pt coming from the anode in spite of the fact that a little Cl^- was added with the Pt. The results show that Pt deposited on top of Cu during electrolysis is only partially dissolved by treatment with 6 N HNO_3 . To test for Cu, diethyldithiocarbamate was used. W. T. H.

New method for determining mercury in mercuric cyanide. E. CATTELAIR *Ann. chim. anal. chim. appl.* 13, 97-8(1931).—See C. A. 25, 1179. W. T. H.

Determination of traces of mercury. ALFRED STOCK, HERMANN LUX, FRIEDRICH CUCUEL and FRANK GERSTNER *Z. anorg. Chem.* 44, 200-6(1931).—After considering the various methods that have been proposed for the detn. of traces of Hg, a method is described which permits the exact detn. of as little as 1×10^{-4} mg. of Hg and the detection of 2×10^{-4} mg. The method depends upon the electrolytic deposition of the Hg on a Cu wire from a soln. contg. $HgCl_2$. After the electrolysis, the wire and its deposit are dried and heated. The Hg distills off and is collected in a capillary. It is finally made to collect together into a tiny globule, the size of which is measured under the microscope. To det. Hg in urine, feces, etc., the Hg is pptd. together with about 20 mg. of added Cu by introducing H_2S into the chlorinated soln., the mixt. of CuS and HgS is transformed to chloride and the soln. (not over 20 cc.) is electrolyzed. The method appears to be extremely sensitive and remarkably accurate but there are so many possible sources of error that the original paper should be studied carefully before attempting to make an analysis. By means of this method, Hg has been detected in nearly every reagent in the lab., especially in HCl which has in every case showed a positive test. Blue vitriol, $KClO_4$, tap water and distd. water are among the few things that have been found to contain no Hg. W. T. H.

Determination of a very small quantity of cadmium in a rich zinc ore. J. G. FAIRCHILD *Chemist-Analyst* 20, No. 3, 5-7(1931).—A soln. 3 N in H_2SO_4 is recommended for the pptn. of CdS in a sample sufficiently large to cause some pptn. of ZnS . The ppt. is dissolved and the pptn. repeated twice more. By evapn. with H_2SO_4 , $CdSO_4$ is obtained and weighed. W. T. H.

Modification of Penny's method for determining iron in ores. ROY P. HUDSON *Chemist-Analyst* 20, No. 3, 6-7(1931).—The titration with diphenylamine as inside indicator is recommended with H_2SO_4 - H_3PO_4 mixt. added, as in $KMnO_4$ titrations. Full directions for analyzing an ore are given. W. T. H.

Detection of the tin group metals and their separation from the copper group by means of ammonium monosulfide. LEO LEHRMAN *J. Chem. Education* 8, 846-51(1931).—By passing H_2S through 10 cc. of concd. NH_4OH for 15 sec. a reagent is obtained which is recommended for the sepn. of the sulfides of As, Sb and Sn from those of Hg, Bi, Pb, Cu, Bi and Cd. The advantage of such a reagent is that it does not dissolve appreciable quantities of the latter group of sulfides and the disadvantage is that it does not dissolve SnS well. The claim is made, however, that enough Sb_2S_3 and SnS dissolve to give pos. tests. W. T. H.

Determination of silver in steel. WALTER BERG *Chem.-Ztg.* 55, 259(1931).—Treat 10-15 g. of drillings with 250 cc. of 6 N HCl , evap. to dryness and take up with HCl as in the detn. of SiO_2 . Filter off the insol. residue, which will contain $AgCl$, and

wash with hot water. Digest the ppt with 5 N NH_4OH and filter through the original filter. Make the ammoniacal soln acid with HCl and set aside in a dark place. Heat the filtrate from the original SiO_2 pptn and introduce a vigorous stream of H_2S for 0.5 hr. Filter, wash with H_2S water, dissolve the ppt in aqua regia and evap. to dryness. Take up with 6 N HCl , make ammoniacal, heat to boiling and filter. Make the filtrate acid with HCl , heat to boiling to coagulate the AgCl and filter off this ppt. together with that originally obtained. Dry and weigh. W. T. H.

Thermal reduction of sulfur in a hydrogen atmosphere applied to the analysis of iron and steel. YOSHITO YAMACHI. *Anniversary Vol. Dedicated to Masumi Chikashige* 1930, 111-8.—The assumption is made that aged samples of iron and steel contain free S and sulfates. The following procedure is shown to take care of these constituents. Weigh out a suitable quantity of sample into a porcelain boat and heat for 30 min at 800° in a stream of H_2 , passing the gas from the furnace into ammoniacal Cd soln. Det. the S in the resulting CdS by the usual iodometric method. Take the ignited residue, transfer to a flask and det. its S content by the usual evolution method for detg. S in steel. W. T. H.

Quantitative spectral analysis of metallic alloys. GASTONE GUZZONI. *Atti III congresso naz. chim. pura applicata* 1930, 636-7.—See C. A. 24, 2601. G. M. MURPHY.

Estimation of ferric oxide and magnetic oxide in the presence of iron and ferrous oxide. E. W. CHEESBROUGH. *Chemist-Analyst* 20, No. 3, 14-5 (1931).—For detg. metallic Fe, treat 1 g. with 100 cc. of 20% CuSO_4 soln, filter, make acid with H_2SO_4 , and titrate with KMnO_4 . To det. FeO , treat the residue from the Fe detn. with 50 cc. of 7 N H_2SO_4 in a stoppered flask, filter, wash and titrate the filter with KMnO_4 . To det. total Fe, proceed in the usual manner. To det. Fe_2O_3 , place 0.5 g. of sample on filter paper, rub the mass with another paper, tap off all powder from the latter, remove Fe_2O_3 with a magnet and weigh. W. T. H.

Volumetric determination of magnesia and alumina in minerals and refractory stones. HORST ECKSTEIN. *Chem.-Ztg.* 55, 227 (1931).—The following procedures are said to be accurate, but no analytical results are shown. To det. Mg in ores, slags, etc., take up 2.5 g. of sample in HCl , remove SiO_2 as usual and ppt. Fe and Al twice with NH_4OH . Make up the filtrate to 500 cc. and take 20 cc. for the Mg detn. Dil. to 100 cc., add 3-5 g. NH_4Cl , 15 cc. of 3% 8-hydroxyquinoline and 2 cc. of concd. NH_4OH . Boil, filter, wash, dissolve the ppt in HCl and titrate with 0.1 N KBr-KBrO_3 . To det. Al in refractory materials, fuse 0.5 g. of the sample in a Pt crucible with 5 g. of borax, take up the melt in dil. HCl and remove SiO_2 in the usual way. Make up the filtrate to 250 cc. and take 25 cc. for further analysis. Dil. to 100 cc., add 5 g. NH_4Cl and NH_4OH until a faint turbidity results. Add an excess of the reagent, heat to 70° and det. the Al in the ppt. with 0.1 N KBr-KBrO_3 soln. W. T. H.

Determination of lime and magnesia in roasted pyrite. LOUIS SCHREIBER. *Ann. chim. anal. chim. appl.* 13, 98-102 (1931).—The method described calls for treatment with concd. HCl , dehydration and removal of SiO_2 , double pptn. with NH_4OH and H_2O_2 , pptn. of CaC_2O_4 and MgNH_4PO_4 . No attempt is made to remove Pb or Cu but, when the former is present, NH_4OAc is added to prevent formation of Pb oxalate or phosphate. No analytical results are given to show that the procedure is accurate. W. T. H.

Potassium permanganate as a reagent for the detection of lower oxides in phosphoric acid. JOHN W. SMITH. *J. Chem. Soc.* 1931, 528-9.—To detect lower oxides in P_2O_5 , dissolve 1 g. of sample in 10 cc. of water in a glass-stoppered test tube. In a similar tube place 10 cc. of pure H_3PO_4 soln. of approx. the same concn. Add 1 cc. of 0.001-0.0002 N KMnO_4 to each soln. and allow to stand for 2 hrs. If there is any perceptible difference, the lower oxide is present. In this way as little as 0.002 mg. of P_2O_5 can be detected. For these small quantities of impurity it is well to keep the tube immersed in water at 60° for 2 min. W. T. H.

Quantitative analysis of phosphoric acid. V. New colorimetric determination. MASAYOSHI ISHIBASHI. *Anniversary Vol. Dedicated to Masumi Chikashige* 1930, 1-7; cf. C. A. 23, 4645.—The method depends upon pptg. the PO_4^{3-} as MnNH_4PO_4 and the colorimetric detn. of Mn as MnO_4^- after oxidation with NaBiO_3 in HNO_3 soln. In the expts. described, in which the P content varied from 4 to 0.2 mg. of $(\text{NH}_4)_2\text{HPO}_4$, the pptn., dissolving and treatment with bismuthate all took place in a centrifugal tube of 15 ml. capacity. To the soln. of $(\text{NH}_4)_2\text{HPO}_4$, a little 10% NH_4OH HCl soln. and an excess of MnCl_2 were added together with about 100 mols. of NH_4Cl for each mol. of MnCl_2 . The soln. was heated to about 90° and concd. NH_4OH added to effect complete pptn. of MnNH_4PO_4 . After 5 hrs. the ppt. was washed with a slightly ammoniacal molal soln. of NH_4Cl contg. a little NH_4OH . The ppt. was dissolved in 4

It seems probable that the Clark and Jelley test depends on some substance other than $S_2O_4^{--}$ which is formed from $S_2O_8^{--}$ by oxidation with I. By titrating dil $S_2O_8^{--}$ solns with I-starch azide mixt the method may be made quant. In the discussion Clark and Jelley showed a new reaction: a soln of NaN_3 and $K_2S_2O_8$ forms $S_2O_4^{--}$ and N_2 on heating. They think it possible that nascent $S_2O_4^{--}$ is sufficiently active to react with N_3^- , re-forming $S_2O_8^{--}$ which then reacts with the excess of I present, re-starting the cycle, so that a small quantity of $S_2O_4^{--}$ can cause decompa of a large amt of azide. L. E. JELLEY

The detection and determination of methanol. LUCIEN SEMICHON and MICHEL FLANZY *Ann fals* 24, 80-7(1931).—A complete oxidation of MeOH to CO_2 and H_2O can be accomplished by excess $K_2Cr_2O_7$ in H_2SO_4 soln, when MeOH is in excess the products consist of a mixt. of CH_3O , HCO_2H and methylal or of a mixt. of HCO_2H and methylal. The methods proposed to date for the oxidation of MeOH to CH_3O give a yield of only 3-10% of theory, but S and F obtain a yield of about 32% by first detg the total oxidizability of the mixt. Details of the technic are described. The CH_3O can be identified by means of NH_4OH , $PhNMe$, or preferably $PhNHNH_2 \cdot HCl$ by the method of Bertrand, for the detn a known amt of MeOH is simultaneously treated in exactly the same way as the sample, and the 2 solns are compared colorimetrically. A. PAPINEAU COUTURE

The de Haas fluorescence reaction of β -naphthol. N. SCHÖDL *Pharm. Woch-*
blad 68, 279-80(1931).—The green fluorescence described by de Haas (*C. A.* 25, 1183), obtained when an aq. soln. of β -naphthol is treated with glacial AcOH and concd. H_2SO_4 is added, does not occur immediately with purified AcOH unless a trace of AcH is added. It does occur, however, with highly purified AcOH if the reaction mixt. is allowed to stand for a long time. A. W. DOX

The Hehner test for formaldehyde (in milk). CHARLES C. FULTON *Ind Eng Chem, Anal Ed* 3, 199-200(1931).—The Hehner test is improved by using Br as oxidizer and by dilg the H_2SO_4 . The improved test is sensitive to 1:1,000,000. Two procedures are given. For a zone reaction dil 5 cc. concd. H_2SO_4 with 1 cc. H_2O and cool. To 3 cc. of this diluted acid add a small crystal of KBr. Shake and then overlay with 1 cc. of the milk to be tested. A violet zone quickly develops. To obtain a uniform color throughout the soln, dil 8 cc. concd. H_2SO_4 with 5 cc. H_2O and cool. To 4 cc. of this diluted acid add 1 cc. of the milk and mix under cooling. Add 0.5 cc. of a mixt. of equal vols. of concd. H_2SO_4 and satd. Br water and shake. A violet color develops at once. This procedure is well suited for distillates. Dil the acid with the distillate to be tested and use pure milk. C. A. ABLEDO

The opium alkaloids as reagents for formaldehyde. CHARLES C. FULTON *Ind Eng Chem, Anal Ed* 3, 200-1(1931); cf *C. A.* 24, 467.—The addn. of an oxidizing agent increases the sensitivity of the test for detecting CH_2O with alkaloids. With apomorphine, codeine or pseudomorphine the test is unmistakably sensitive to 1:500,000, with papaverine, to 1:100,000. The oxidizing agent used is a 10% soln. of $Fe_2(SO_4)_3$, or 5 drops of concd. HNO_3 in 50 cc. water. The tests are not sp. for CH_2O but when carried out on a distillate its identification is fairly certain. Papaverine gives the most nearly sp. test. C. A. ABLEDO

The iodometric determination of copper and its use in the estimation of reducing sugars. R. INTONTI *Ann chim applicata* 20, 583-90(1930).—The iodometric titration of Cu_2O dissolved in HNO_3 is preferred to the result obtained by the vol. procedure of Fehling. To make sure that no HNO_3 is present, treatment with $CO(NH_2)_2$ is recommended. A. W. CONTIERI

Bromopotentiometric titration of β -naphtholsulfonic acids in presence of each other. III. Mixtures containing oxy-Tobias acid, F acid and trisulfo acid. J. SMITH HARLAND, STANLEY D. FORRESTER and DAVID BAIRN *J. Soc. Chem Ind* 50, 100-2T (1931); cf *C. A.* 25, 474.—The β -naphtholsulfonic acids, with no sulfonic group in the 8-position, can be detd. in the presence of those acids which have the SO_3H group in this place, but an excess of KBr is necessary for the potentiometric titration if crocein or oxy-Tobias salts are present. In the presence of 10 N H_2SO_4 , the OH group of the β -naphtholsulfonic acids is converted completely to sulfate and this, when associated with the SO_3H group in the 8-position, prevents the Br from entering at the 1-position. F acid exists as the tetra-, di-, and mono-hydrates as well as in the anhydrous condition. The tetrahydrate m 67°, the dihydrate m 95°, the monohydrate m 103-9° and the anhydrous acid m 115-6°. The method recommended for analyzing the mixts. is that recommended in Part II. W. T. H.

Qualitative testing of acetic acid. E. TSCHIRCH. *Oesterr. Chem-Ztg* 34, 38-40 (1931).—The red color produced with $FeCl_3$ is not a reliable test for acetate, the for-

mation of ethyl acetate, to be identified by the odor, is a useful test but depends too much upon the personal equation. The caecodyl test is not at all reliable as the higher homologs of AcOH give a similar result. The test with phthalaldehyde and NH_4OH , which gives a bluish green color, is less well known, it is fairly sensitive but is given by the homologs of AcOH. The best test is that of treating 1-3 cc. of soln successively with 1 cc. of 5% $\text{La}(\text{NO}_3)_3$ soln., 1 cc. of 0.02 N I soln. and a few drops of NH_4OH . On gradually heating to boiling, a starch blue color develops, which is a good test not only for La but also for Ac^{1-} . The homologs of AcOH give similar color effects and SO_4^{--} or any anion that ppt. La must be removed. The fact that AcOH can be dried off easily from a soln. often helps in the prepn. of suitable solns. for the test. SO_4^{--} and PO_4^{--} can be removed by pptn. with Ba and when large quantities of NO_3^- or Cl^- are present, advantage can be taken of the fact that 100 cc. of abs. alc. dissolves 1.81 g. of NaOAc but only 36 mg. of NaNO_3 and 65 mg. of NaCl . W. T. H.

Methods for the analysis of technical solvents. II. Color tests for benzene, nitrobenzene, toluene and xylene. HANS H. WERNER. *Chem. Ztg.* 55, 201-3 (1931).
 C A 24, 20-3.—For the colorimetric detection of these aromatic hydrocarbons it is well to transform them first into polynitro compds. Take not over 0.5 cc. of the solvent, or more if it is a soln. in gasoline or the like, and mix with 1 cc. of concd. HNO_3 and 2 cc. of 100% H_2SO_4 . Shake 1 min. and cool. Take 0.1 cc. of the acid layer and dil. with 1 cc. of water, which will cause the formation of a suspension in water of the nitro compds. Sufficient water must be used so that the alc. subsequently added will not be attacked. *Detection of benzene*.—Mix 0.5 cc. of the aq. suspension with 1 cc. of isoamyl alc., add 2 N NaOH till alk. and shake for 1 min. with a stick of NaOH 1 cm. long. If only benzene is present, the amyl alc. layer remains practically colorless, but if the higher homologs are present a dirty brown paste is formed with fugitive colorations. Now, with still a little solid NaOH remaining undissolved, add 1 cc. of acetone, an intense blue to permanganate-violet color should form, becoming brown on long standing. By this test benzene can be identified in the presence of its homologs and in mixts. with I oil or $\text{C}_{11}\text{H}_{14}$. *Detection of nitrobenzene*.—This gives the same test as benzene, but on account of the great difference in b. p. as well as odor, there is little danger of confusing the two. In case of doubt, however, take 10 cc. of the original sample and treat in a large test tube with 10 cc. of abs. alc. and 5 cc. of concd. HCl , add a small piece of Sn and some coppered Zn , heat 15 min. on the water bath, sep. from undissolved metal in a separatory funnel, pour off the liquid from the undissolved metal into a small separatory funnel, draw off the bottom layer, add a little KClO_4 and 0.5 cc. of concd. HNO_3 (Pezet's test for aniline), after the mixt. has stood without shaking, rings with red and blue edges form if only a trace of nitrobenzene was originally present (reun. oils give brownish red colors and some phenols give pale red effects). *Detection of toluene*.—Mix 0.5 cc. of the aq. suspension with 1 cc. of pure benzyl alc., add 2 N NaOH to alk. reaction and shake. The alc. layer will be colored brownish red, a violet shade sometimes appearing first. *Detection of xylene*.—Mix 0.5 cc. of the aq. suspension with 1 cc. of cyclohexanol, make alk. with 2 N NaOH and shake well. The alc. layer should be an intense green. This test depends upon a reaction of m -xylene, which constitutes 70-85% of com. xylene. It can be used successfully in the presence of benzene and toluene when not more than 10% of these lower homologs is present in the original solvent. In case of doubt try the test after adding 95% alc. to the original solvent (xylene fraction) and then water until a turbidity results, xylene will sep. before benzene or toluene, and in this way a portion rich in xylene can be obtained and the presence of 2.5% xylene detected in a mixt. W. T. H.

Furfural as an industrial poison and its determination in air (KORENMAN, REYNIE)
 13. Organic Pb compounds [determination of Pb in aryl Pb compounds] (AUSTIN)
 10. Quantitative chemical analysis by means of the absorption of x rays (MOORE)
 3. Apparatus and methods for precise fractional distillation. Analysis. New method of gas analysis (PODBIELNIAK) 2.

Analytical Reagents. Standards and Tests. 3rd ed., rewritten. Compiled first by EDWARD WHITE. London: Hopkin and Williams, Ltd. 135 pp. 2s. 6d. Reviewed in *Pharm. J.* 126, 364 (1931).

CHARLES AND MARTIN. Recueil de manipulations de chimie et métallurgie. Fasc. I. Analyse chimique qualitative et quantitative. Paris: Vuibert. 96 pp. F 10.

MICHELIS, L. Éléments de chimie analytique. 2 vols. Louvain: Uystpruyt. F. 40 per vol.

Separating arsenic from vanadium and phosphorus. Oskar JOCHIM Ger. 521,191, June 10, 1929. A soln contg these elements as alkali or NH_4 arsenates, vanadates and phosphates is sat'd at 30–100° with SO_2 or a bisulfite, and the soln is then boiled to remove excess of SO_2 . A ppt contg the V and P but no As is thus obtained.

8—MINERALOGICAL AND GEOLOGICAL CHEMISTRY

EDGAR T. WHERRY AND J. P. SCHAIFFER

Practice of mineral polishing. WERNER GRÜNDLER *Metall u. Erz* 28, 143 7 (1931).—Various methods of prep'g specimens for microscopic examn are investigated. Relief polishing is not always the most suitable means of bringing out the structure of the sample. Plane polishing in which the structure is brought out in only very weak relief, is often more satisfactory. An oil emery mixt is found suitable for this purpose. The details of mounting and prep'g the specimens are given. H. STÖRZT

Stephanite, argentite and silver, South Lorrain, Ontario. T. L. WALKER Univ. Toronto Studies, *Geol. Ser.*, No 29, Contributions Can. Mineral, 1930, 13–5.—The Ag minerals occur in veins with a gang of calcite assoc. with arsenides of Co, Ni and Fe. Some vein cavities contained argentite, wire Ag, stephanite, proustite and rarely pyrostilpnite followed by a final deposition of calcite lining the cavity walls. Argentite was the most abundant. J. W. SHIPLEY

Mineralogical study of the marcasite group. FLORA THOMSON Univ. Toronto Studies, *Geol. Ser.* No 29, Contributions Can. Mineral, 1930, 75–83.—A description is given of marcasite, loellingite, arsenopyrite, safflorite, rammelsbergite, glaucodite, alloclaste and nickel skutterudite based on a study of polished mineral sections etched with various special reagents. The phys. and microchem. properties are listed in a table. J. W. SHIPLEY

Fibrous marcasite in crystalline calcite near Logansport, Indiana. FREDERICK R. SMITH AND RUSSELL A. SCHAEFFER *Proc. Indiana Acad. Sci.* 38, 231 (1929).—Analytical data are given to show that a fibrous greenish gold, metallic mineral zonally distributed through the outer portions of calcite masses has the compn FeS_2 . The color of the sulfide indicated it was marcasite rather than pyrite, and this is considered to have been confirmed by soln in conc'd HNO_3 , free S being formed. H. M. S.

Lodestone from Bon Accord, Transvaal. T. L. WALKER Univ. Toronto Studies, *Geol. Ser.*, No 29, Contributions Can. Mineral, 1930, 17–9.—The analysis of a sample from a specimen of lodestone secured 10 miles north of Pretoria, S. Africa, indicated magnetite 54.06%, limonite 9.35%, ferric titanate 35.32%, Si and Mg 1.02% when the Ti was considered as occurring as the dioxide and the reducing action toward KMnO_4 ascribed entirely to Fe^{++} . Since 95% of the specimen appeared to be one mineral species, W. suggests that the Ti is present as Ti_2O_3 , which on going into soln reacts with the Fe_2O_3 as follows: $\text{Fe}_2\text{O}_3 + \text{Ti}_2\text{O}_3 \rightarrow 2\text{FeO} + 2\text{TiO}_2$. A recast of the analysis would give 13.87% for spinel, 76.67% of sesquioxides largely ferric, and the residue limonite. The dominant mineral would then be titaniferous magnetite, $(\text{Fe}, \text{Ti})_2\text{O}_3$. J. W. SHIPLEY

X-ray distinctions between magnesite and dolomite. FRANZ HALLA, *Monatsh.* 57, 1–8 (1931).—Powder spectrograms of very pure samples of the minerals from various sources have been utilized to establish their structure. They can be distinguished with certainty in this manner. The parameters of the rhombohedral unit cell of $\text{CaMg}(\text{CO}_3)_2$ are $a = 6.050 \pm 0.001$ A. U., and $\alpha = 46^\circ 51'$. Detailed studies and reproductions are given. GREGG M. PIVAN

The optical properties of manganese-poor grunerite and cummingtonites compared with those of manganiferous members. N. SUNDIUS *Am. J. Sci.* [6], 21, 330–44 (1931).—Full optical data are given. ALDEN H. EMERY

Pyroxene and scapolite from Templeton Township, Quebec. A. L. PARSONS Univ. Toronto Studies, *Geol. Ser.* No 29, Contributions Can. Mineral, 1930, 25–8.—A description of crystals of pyroxene and scapolite is given together with a chem. analysis of each. J. W. SHIPLEY

Chemical and optical study of amphibole. A. L. PARSONS Univ. Toronto Studies, *Geol. Ser.* No 29, Contributions Can. Mineral, 1930, 29–37.—A crystallographic and chem. study of black crystd hornblende from Montezuma Township, Hastings County, Ontario, was made. An increase in the ws of the amphiboles, closely related to the total Ti and Fe^{+++} oxides, was observed. J. W. SHIPLEY

Calamine, galena, magnetite. T. L. WALKER AND A. L. PARSONS Univ. Toronto

Studies *Geol Ser No 29*, Contributions Can Mineral. 1930, 21-3—Specimens of calcum from the Broken Hill Mines, Northern Rhodesia, were examd crystallographically the form (052) was observed. Galena from the Ivanhoe Pb-Ag mines at Samson B C was found to contain free S. The alteration products form a series of fine parallel bands in the mineral with an occasional small cavity containing brilliant crystals of anglesite. Large crystals up to 10 lbs of magnetite were obtained from Faraday Township Ont. The largest crystal encloses both apatite and lepidomelane and a cavity from which some mineral possibly calcite or apatite, has disappeared. Analysis indicated an almost perfect balance between RO and R_2O_3 . J W S

Chemical formulas of hallynite and cancrinite. L H BOASTROM *Bull comm geol Finlande* 1930, No 92 51-7—An English version of the article abstracted in C A 24, 3196 H C DUUS

Lattice dimensions of heulandite from Wasson's Bluff, Nova Scotia. A L PARSONS *Univ of Toronto Studies, Geol Ser No 29*, Contributions Can Mineral 1930, 37-8—The dimensions of the unit cell for the heulandite were found to be $a = 7.5139$, $b = 17.009$, $c = 15.913$ A and $\beta = 83^\circ 34'$ (Dana). The sp gr of the crystal was 2.22. Calcul from the chem analysis of heulandite from the traps of Digby Gut, Nova Scotia, failed to give a simple formula for the mineral but corresponded to a mixt of a hydrous metasilicate with a hydrous orthosilicate in the proportions of 4:1. From this assumption $N = 1.0113$. Other mixts are equally possible but the evidence indicates an isomorphous mixture of 2 types of mols in heulandite. J W SIMPSON

Lattice dimensions of natrolite from Wasson's Bluff, Nova Scotia. A L PARSONS *Univ of Toronto Studies, Geol Ser No 29*, Contributions Can Mineral 1930, 35-6—The dimensions of the unit cell for the natrolite were found to be $a = 18.306$, $b = 18.708$ and $c = 8.8814$ A. The sp gr of the crystal was 2.257 and there were 8 mols in the unit cell, by using Brögger's value for N , N is found to be 8.0934. J W S

Collophanite from Miocene brown shales of California. E WAYNE GALLIKER *Bull Am Assoc Petroleum Geol* 15, 257-69 (1931)—An impure, pyritized oolitic collophanite occurs in the Miocene brown shales of the southern San Joaquin Valley, Calif. The grains are 0.1-0.5 mm in diam. Their chem compn and phys properties depend on the degree of pyritization. The grains are embedded in a matrix which is dominantly clay. These oolites are syngenetic. The mechanism of pyritization is discussed. ALDEN H EMERY

British Columbia in 1930. ANON *Mining J* (London) *Ann Rev No* (Jan 24, 1931) 25. Eastern Canada in 1930. ANON *Ibid* 21. Gold Mining in Nova Scotia. ANON *Ibid* 27. Norway in 1930. ANON *Ibid* 28. Mining in Newfoundland. ANON *Ibid* 29. The mineral resources of Tanganyika. ANON *Ibid* 31. Nigeria in 1930. ANON *Ibid* 33. The Federated Malay States in 1930. ANON *Ibid* 35. Australia in 1930. ANON *Ibid* 36. Portugal in 1930. ANON *Ibid* 36. Cornwall and Devon in 1930. ANON *Ibid* 37. The Minerals of the Hyderabad State, India. ANON *Ibid* 48-9. ALDEN H EMERY

Quinquennial review of the mineral production of India for the years 1924 to 1928. Chromite. E H PASCOE *Records Geol Surv India* 64, 23-31 (1930), cf C A 24, 1824—Indian chromite is usually assoc with serpentine. In Bihar and Orissa it occurs in bed like veins and as scattered granules in serpentinized saxonites and dunites forming laccolithic intrusions. As in Baluchistan the chromite is of primary (magmatic) origin and contemporaneous with the peridotites. The subsequent serpentinization of the peridotites has been accompanied by widespread silicification with the production of marginal zones of chert. The output is steadily increasing. Graphite. E L C. CLEGG *Ibid* 106-9—Graphite occurs in various parts of India in schists. Magnealite. E H PASCOE *Ibid* 166-72—In the Salem district a network of magnetite veins pierces 2 great intrusive masses of serpentinized ultra basic rocks. Production is steadily increasing. Cyanite. J A DUNN *Ibid* 403-5—Important deposits of cyanite and of cyanite quartz rock occur in India. Sillimanite. J A DUNN *Ibid* 426-9—Sillimanite-coriundum deposits assoc with cordierite biotite quartz microcline and sillimanite quartz schists occur in Assam. ALDEN H EMERY

Relation of the ore deposits of the southern Rocky Mountain region in the Colorado Plateau. B S BUTLER *Proc Colo Sci Soc* 12, No 2, 23-36 (1929)—Ore deposits of the southern Rocky Mountain region are coned in a narrow belt surrounding Colorado Plateau, assoc with volcanic fields which are also largely confined to the same belt. Igneous activity resulted in deposition of the ores. A bibliography is given. E I S. Geological and metallogenetic study of South Africa. A DAMAY *Rev ind minérale* 229, 271-93 (1930)—A review of the essential elements of geology, stratigraphy and eruptive phenomena of South Africa, central Rhodesia and regions further north,

including the province of Katanga of Belgian Congo. Notes are given on the principal deposits of this region which is one of the richest mining districts of the world.

E. I. S.

Impressions of the mineral industry of British South Africa. W. B. TIMM. Can. Dept. Mines, Mines Branch, *Memorandum Series* No. 46, 27 pp (1931).—T. discusses diamonds, asbestos, Pb, Zn, V etc.

ALDEN H. EMERY

Geology of South Park Section. A. LAKES. *Mountain States Mineral Age* 1, No. 3, 7, 20-1 (1930).—Notes on Au, Ag, Pb and coal deposits are given and oil possibilities discussed.

E. I. S.

Geology and ore deposits of Bannack and Argenta, Montana. PHILIP J. SIMMON. *Mont. Bur. Mines and Geol. Bull.* No. 6 (1931).—These 2 districts were the first important placer and lode mines in Montana and have produced Au, Ag, Pb, Zn and Cu. The ore deposits are closely related to intrusive rocks, but occur also in sedimentary rocks near intrusive contacts. Those with obvious exposures have been largely exhausted, but it is possible that additional discoveries may be made.

CURTIS L. WILSON

Geology of the Shonia Lake area, district of Kenora (Patricia portion). H. C. LAIRD. *Ann. Rept. Ont. Dept. Mines* 39, Pt. 3, 1-21 (1931).—Extrusive rocks (greenstones) forming the basement schist series are surrounded by granite. Gabbro-diorites and amphibolites of pre granite age form 40% of the schist area. Descriptions of the rocks and Au prospects are given.

ALDEN H. EMERY

Geology of the area from Minaki to Sydney Lake, District of Kenora. D. R. DERRY. *Ann. Rept. Ont. Dept. Mines* 39, Pt. 3, 25-41 (1931).—The consolidated rocks (all Pre Cambrian) fall into 2 series: (1) a sedimentary-volcanic series, much folded and metamorphosed, and (2) an intrusive series ranging from dioritic to granite and pegmatite. The intrusives occupy 80% of the area. They are described in detail. Mineralization (usually replacement of the country rock) was closely associated with the later phases of intrusion. Quartz veins are rare. Pyrrhotite with minor pyrite and molybdenite carrying Au and Ag is known in the coarser gneiss. Pyrite, in a siliceous groundmass with varying proportions of chalcopyrite and carrying Au, occurs in the volcanics and fine quartzitic sediments. A cobaltite vein has been prospected.

ALDEN H. EMERY

Geology of the Shoal Lake (west) area, District of Kenora. LEONARD GREER. *Ann. Rept. Ont. Dept. Mines* 39, Pt. 3, 43-56 (1931).—A thick and intricate Keweenaw section is exposed. Lava flows, agglomerates and obscure sediments are interbedded with no particular order. Intrusive into them are bodies of granite, felsite and quartz porphyry. Au-bearing quartz veins occur chiefly in greenstone. The rocks are described in detail.

ALDEN H. EMERY

Geology of the Bigstone Bay area, Lake of the Woods, district of Kenora. G. G. SURREL. *Ann. Rept. Ont. Dept. Mines* 39, Pt. 3, 57-71 (1931).—A basic and siliceous series of volcanic rocks, largely extrusive, are intruded by a series of dioritic to gabbroic bodies. All of these are intruded by granite and many small granite, felsite and quartz porphyry dikes. Much later the complex was intruded by narrow diabase dikes. Au-bearing quartz veins occur associated with the contact between the granite and basic schists. Ag, pyrite, arsenopyrite, chalcopyrite, sphalerite, galena and tellurides have been reported. The rocks are described.

ALDEN H. EMERY

Geology of the Eagle-Circle District, Alaska. J. B. MERITT, JR. *U. S. Geol. Survey Bull.* No. 816, 166 pp (1930).—M. discusses the geologic history, distribution of placers, sources of Au and placer mining operations. Two groups of claims, comprising 7 mi. of placer ground on Coal Creek, are regarded as good hydraulic venture.

E. I. S.

Geology of the Panamint Silver District, California. F. MACMURPHY. *Econ. Geol.* 25, 305-25 (1930).—The Ag bearing fissure veins occur largely in limestone. They comprise: quartz-galena-pyrite veins and pyrrhotite-pyrite-quartz veins. The mineralogy of ores is described and origin and age of deposits discussed. A bibliography is given.

E. I. S.

Mineral association at the Marble Bay Mine, Texada Island, B. C. T. L. WALKER. *Univ. Toronto Studies, Geol. Ser.* No. 29, Contributions Can. Mineral, 1930, 5-8.—The ores of the Marble Bay mine are chalcopyrite and bornite, rarely native Ag and molybdenite. The non-metals are represented by garnet, wollastonite and epidote. The deposits occur at the contact of a dioritic porphyritic intrusion with Mesozoic limestone. Analyses of the bornite, wollastonite, andradite and grossularite are given.

J. W. SHIPLEY

Geological and microscopical study of some copper deposits of China. C. Y.

Hsien *Bull Geol Soc China* 8, No 4, 263-266(1929)(in English) — Deposits in Hsien Yunnan and Szechuan are described. E I S

Lubi and Lukula lead-copper deposits, their origin and their relation to stratigraphy of the region. I ICHIKAWA *Ann soc geol Belg* 52, No 11, 85 pp (1929) — N'Tendu ore occurrence is regarded as float mineral and not a deposit in place. A discussion of the associated minerals and a theory of the origin of the deposits are given. E I S

Iron and manganese deposits of Brazil. L P DE OLIVEIRA *Linha mines et metal* 57, 626-7(1929) — Iron deposits constitute the largest mineral resource of Brazil. Outline descriptions of principal Fe and Mn ore deposits are given, with analyses of Mn ore. E I S

The iron ore deposits of South Africa. HANS SCHNIDERKNECHT *Arch Eisen hüttenw* 4, 269-76(1930) A survey of the nature, geographical positions and sizes of the Fe ore deposits of So Africa based on the latest literature and some personal observations. J BALOGH

World deposits of nickel. JI CARLBORF *Jernkonkorets Ann* 1929, 493-4 — C gives descriptions of Ni ore minerals, mining and refining processes, and mine production statistics for 1913 to 1927. The present yearly consumption is 40 000 tons. A bibliography is appended. E I S

Platinum deposits of South Africa. P KILICK *Glückauf* 66, 868-74(1930) — A review. E I S

Tin deposits of Morococals, Bolivia. R GIERER *Z prakt Geol* 38, 113-21, 135-42(1930) — A geological treatment, including description of the deposits. E I S

Mercury deposit in Coso range, Inyo County, California. T WAGNER *Mtn in Calif Calif Div of Mines and Mir* 26, No 1 29 63(1929) — In a crater like opening in a conical hill of altered rhyolite, there is a network of veins, seams, crevices and fractures through which cinnabar crystals and free S have been deposited. Average samples contained 0.16 and 0.19% Hg. Selected ore 7 to 11 percent. The origin of the deposit is discussed. E I S

The geology of gypsum deposits. V CHARRIN *Rev matériaux construction trav public* 1931, 99-101 — The deposits of gypsum near Paris and in Lorraine or the Provençales occur either in Triassic or Miocene formations. F O ANDRÉ

Gypsum of Ping Lu District, South Shensi. S I TAO *Bull Geol Soc China* 8, No 4, 327-41(1929)(in English) — Veins, veinlets and beds of gypsum occur in red marl both conformable and oblique to bedding. It is mined in a primitive manner, the output being about 10 tons per day and price at river bank \$5 40 per ton. E I S

Alabaster deposits at Campredon near Carcès. V CHARRIN *Géologie civil* 97, No 1, 18-9(1930) — Geological features of deposits and methods of mining are described. E I S

Kentucky fluorites. W R JILSON *Pon Am Geol* 54, No 1, 29-30(1930) — Production data, 1926 to 1929 are given. E I S

The origin of the Lushienburg diamonds. ALFRED WILLIAMS *S African Min ing and Eng J* 41, Pt 1, 314(1930) — Geological. ALDEN H EMERY

Carbons and industrial diamonds in 1930. I M VAN MOPPEL *Mining J Ann* Rrr No 9 Jan 24, 1931. ALDEN H EMERY

Contribution to study of evolution of fuels of the coal basin of Liège. M LÉGERAT *Rev min mines* [8], 4, No 1, 5-11(1930) — The nature of coal and variations in content of volatile matter are discussed. A map is shown of distribution of coal and coal formations in the Liège basin, Campine, Limbourg and western Germany. Modifications are attributed to tectonic forces. E I S

Genetic connection between salt deposits and petroleum. A WEINZELD *Intern 2. Bohrtech, Erdölbergbau u. Geol*, May 1, Oil and Gas J 28, No 5, 159(1929) — Contrary to the current idea of migration of oil from mother rocks, W assumes that the greater part of the oil found in porous beds is primary in character, formed in rocks where actually found by drill. His explanation is that animal organisms with soft bodies were abundant where the NaCl content of the water was high, and that these organisms constituted the mother substance of petroleum. E I S

Some factors of solidification in relation to metallurgical and geological problems. S W SMITH *J. Chem Met Mining Soc S Africa* 31, 19-8(1931), cf C A 24, 3970, 5004, 5675, 25, 476, 1184. ALDEN H EMERY

Separation of solid phases in heavy liquids and melts with aid of the centrifuge. W JAEGER *Glückauf* 66, 661-8(1930) — J discusses applications of the centrifuge to petrography, giving new methods of separ with its aid, gravimetric and volumetric processes, with aid of heavy liquids, and details of Wolff centrifuge. E I S

Petrography of the rocks in the vicinity of Killarney, Ontario. W. A. JONES. Univ. of Toronto Studies, *Geol. Ser.* No. 29, Contributions Can. Mineral 1930, 39-60.—A detailed study of the rock types in the area adjacent to Killarney Village, Ont., was made. A sharp contact was observed in the field between the coarse-grained Killarney granite-gneiss on which the village stands and the fine-grained granite-gneiss to the east. Dikes of the fine-grained rock cut the coarse-grained. At the contact of these 2 rock types the fine-grained granite-gneiss has apparently been chilled, since its texture is unusually fine. A series of chem. analyses of samples taken across the strike of the rock from the quartzite into the coarse-grained rock showed no progressive decrease in silica and increase in alkalis. Microscopic examn. of thin sections did not reveal proof of the origin of the fine-grained granite-gneiss by the metasomatic replacement of quartz by feldspar. The Killarney granite-gneiss is intrusive into sediments of Huronian age. Microscopically and chemically, its character is that of an igneous rock of plutonic origin. The compn. of the coarse-grained rock is not widely different from that of the fine-grained, so that the latter is probably a slightly younger, more quickly cooled intrusive rock which has originated in the same magma reservoir. J. W. S.

Certain xenoliths occurring in gabbro at Sudbury, Ontario. W. A. JONES. Univ. of Toronto Studies, *Geol. Ser.*, No. 29, Contributions Can. Mineral 1930, 61-73.—The inclusions of white rock at Sudbury originated contemporaneously with the enclosing gabbro as a result of the action of the magma on a nucleus of siliceous sedimentary material. No remnants of foreign material were found to prove this conclusively. Actual fusion of the inclusions did not occur. The process was one involving the penetration of materials from the magma which reacted with quartzite to form hornblende and plagioclase in an essentially solid xenolith. Mol. transfusion of material from the magma into the xenolith accounts for the mineralogical symmetry of the inclusion. The penetrative ability of Fe and Mg would not be as great as Al, Ca, Na and K, and hence the border zone of the xenolith would tend to become rich in Fe and Mg and the formation of hornblende crystals would result. At the edge of the inclusion crystal centers would be close together and a medium textured aggregate made up chiefly of hornblende would result. Within the xenolith the penetration of Mg and Fe would not be so great, with the result that large crystals would gradually develop. Impoverishment of the reacting liquid in Mg and Fe due to the pptn. of hornblende would result in the development of bytownite in the interstices between the hornblende crystals. It is thought that Na and K would explain the presence of andesine in the center of the body rather than bytownite. The presence of a great no. of fluid inclusions in the residual quartz of the xenoliths indicates that the reaction between magma and quartzite took place in the presence of mineralizers. J. W. SHIPLEY.

The aqueous chilling of basaltic lava on the Columbia River Plateau. RICHARD E. FULLER. *Am. J. Sci.* [5], 21, 281-300(1931).—A petrographic study. A. H. E.

Dalmatianite, the spotted greenstone from the Amulet Mine, Noranda, Que. T. L. WALKER. Univ. Toronto Studies, *Geol. Ser.* No. 29, Contributions Can. Mineral 1930, 9-12.—Microscopic and chem. examn. of the greenstone indicated considerable variation in the compn. of the specimens, some being fairly acid while others are quite basic. The S and Zn contents indicate a change from the original volcanic type due to contact with the sulfide ore body. The spots represent amygdaloidal filling. J. W. SHIPLEY.

Compaction of sediments. PARKER D. TRASK. *Bull. Am. Assoc. Petroleum Geol.* 15, 271-6(1931).—The initial H₂O content of sediments varies with the fineness of the constituent particles. A tentative estimate indicates that it is approx. 45% in well-sorted fine-grained sands, 60% in silts, 80% in clays and > 90% in colloids. A. H. E.

The influence of withdrawn chemical substances in creating empty underground spaces and pressure. CESARE VERCELLI. *Ind. mineraria* 5, 81-2(1931).—Very strong pressure noted in the Monte Cam pits was studied as to its cause, with the conclusion that the ZnCO₃ and PbCO₃ of these were produced by the oxidation of the ZnS and PbS through the action of water, CaCO₃ and MgCO₃, and that the last 2 were transformed into the sol. sulfates and were carried off by water, leaving empty spaces, while the ZnS and PbS remained as insol. carbonates. R. SANSONE.

Chemical transformations caused by the decomposition of vegetable wastes. P. VINASSA DE REGNI. *Ind. mineraria* 5, 21(1931).—Many plants, or plant parts, die every season, and the waste produced contains mineral salts and org. matter with H, C, O, N. If in contact with much O as in tropical regions, a slow combustion follows, while under woods or in immature peat bogs in large masses, contact with atm. O is difficult, and these become putrid by the action of bacteria and fungi, and H₂O, CO₂ and a little N are produced, there being only as exceptions solid remains as lignobolites. With excess

of H₂O and absence of O₂ or if the mass freezes, or the ground is calcareous or has salt, the fungi and decomposing bacteria do not vegetate, or even die. Besides this, if in a year longer and bacteria cannot act on all the vegetable mass, this is increased by other additions, and the wastes can rise for several yards being dry at the top and wet and cottony at the bottom. There is then formed moist material where humus predominates. This is a collection of negatively charged colloids which remain such as long as there is water, resisting flocculation by univalent ions while they are more sensitive to bivalent such as Ca. Acid humus reacts in an opposite way to O₂ and reduces and renders sol ferrous compounds. In layers under sand and peat bogs all the Fe and CaO is eliminated in this way.

Recent advances in science geology. C. W. TIRRELL. *Science Progress* 25, 574-86 (1931). — A review of recent work on the geochemistry of the igneous rocks.

JOSEPH S. HARRISON

Modern technique in the investigation of opaque minerals (PILLER) 9. Pneumatolytic synthesis of silicates (VAM NIEUWSTADT, KILWEAPAL) 6. The Charleston coalfield (PILLER) 21. Refracting minerals from Great Bear Lake, Northwest Territories (SPENCE) 3. Discoloration of minerals by β and γ rays (HICKMAN) 3. Cretaceous limestone as petroleum source rock in northwestern Venezuela (HEDBERG) 22. III (TOLAN) 18.

Fortschritte der Mineralogie, Kristallographie und Petrographie. Band XV, Teil 1. Edited by W. LITTL. Jena: Gustav Fischer, 72 pp.
SCHMIDT, LEONOLD. Der Bernstein. Separate of "Handbuch der Mineralchemie." Dresden: T. Steinkopff, 104 pp. M. 7.

9—METALLURGY AND METALLOGRAPHY

D. J. DEMAREST, R. W. GILLET AND RICHARD KIMBACH

The fiftieth anniversary of Rudolf Vondráček. FERDINAND SETTLER. *Chem. Listy* 25, 73-5 (1931). — A review of V's contributions to physicochemical analysis of metallurgy.

FRANK MAREST

Bibliography of the metallurgical work of the U. S. Bureau of Mines in 1930. R. S. DEAN. *Bur. Mines Information Circ.* 6449, 7 pp (1931). — Over 60 references are given to articles on fundamentals, ore dressing, Fe and steel, hydrometallurgy, Mn investigations, rare and precious metals, non-ferrous metals, lab. app., methods and discussions.

The metallurgy of some of the rarer metals. C. J. SMITH. *Metal Ind.* (London) 38, 836-8 (1931). — The development of electric lamps and thermionic valves has created a demand for metals and alloys with very special properties. While W has generally met the requirements for lamp filaments, Mo, Ta and Al are used for electrodes of thermionic valves. The production of ductile Zr and of ductile Ta and of Mo are outlined as is the manufacture of sheet Mo.

W. H. BOYNTON

Calculation of metal yield in ore-dressing processes. OTTO SYRÉN. *Mettall u. Erz* 28, 140-3 (1931). — Methods of calc. yields in the various ore-dressing processes are discussed.

H. STORRE

The scientific fundamentals of colloidal ore dressing. CARL GOETZ AND ALEXANDER HOROVIC. *Intern. Bergbau u. Hütten* 24, 37-42 (1931). — The elementary theoretical principles relating to suspensions are reviewed. The peptization of a complex Rammelsberg ore analyzing SiO_2 19.67, S 24.63, CaO 16.31, Cu 8.3, Pb 5.17, Fe 17.28 and Zn 5.28%, and sieved through 10,000 mesh was studied by shaking with citric acid and soda for definite periods, settling, and dec. the quantity in suspension. The max. quantity of ore remaining in colloidal suspension depends on the citric acid concn. and on the ratio of ore to peptizing soda. The results are in agreement with Ostwald's peptizing rule.

ARTHUR FLEISCHER

Modern technique in the investigation of opaque minerals and ores. F. COLAS PILLIER. *Science Progress* 25, 633-41 (1931). — A review with sections devoted to (a) the prep. of polished surfaces, (b) reflecting power, (c) anisotropism, (d) other properties and (e) a bibliography of 24 references.

JOSEPH S. HARRISON

Flotation tests on converter slag. FRANK S. WARTMAN. *Bur. Mines, Rept. of Investigations* 3068, 7 pp (1931). — It is possible to recover by flotation about 60% of the Cu content of a chilled converter slag contg. 3% Cu and to obtain flotation concentrates

contg 35% Cu Reagents were 0 14 lb No 5 G N S, pine oil per ton of slag, 0 1-0 15 lb amyl xanthate, 1 25-3 0 lb H_2SO_4 and 0 12-0 25 lb P E oil A H F.

Influence of the iron content of zinc blende upon its adaptability for flotation. LUDWIG KRAIBER *Metall u Erz* 28, 128-130 (1931) See C A 25, 2082 H S.

Operating practice in the copper-flotation plant of Minas do Valle do Vouga, Portugal. G TELFER *Metall u Erz* 28, 125-8 (1931) The installation at this plant is described The ores are complex (Cu pyrites, Zn blende, Fe pyrites) The processes used are described, and the difficulties and methods of overcoming them are recounted H STOFRTZ

Leaching copper from its ores. JOHN D SULLIVAN *J Chem Education* 8, 829-47 (1931) REVIEW C H

Extraction of soluble copper from ores in leaching by percolation. JOHN D SULLIVAN AND KENNETH O BAYARD *Bur Mines, Rept of Investigations* 3073, 43 pp (1931)

The phys nature of the rock its porosity degree of kaolinization and sericitization, and the amt of natural slime det the amt of soln that will be retained and the difficulty that will be found in washing out the sol salts The rate of removal of sol Cu decreased with increases in the amt of slime present and increased only slightly with temp Wet charging the ore into the vat was advantageous In washing, the cycles of submergence and draining should be as close to each other as possible, consistent with fairly thorough mixing of the $CuSO_4$ in the wash soln and thorough draining of the soln The $CuSO_4$ concn was greater at the bottom than at the top of the ore column Circulating the solns increased extn rates Piston displacement washing removed Cu more rapidly than batch washing, but the vol of wash soln required was greater Downward and upward percolation of wash solns gave virtually identical extn rates Less $CuSO_4$ was added in the ore when the impregnating solns were acidic ALDEN H EMERY

The extraction of copper from oxidized ore by cyanide solution. E T DUNSTAN *J Chem Met Mining Soc S Africa* 31, 190-4 (1931) Cyanide extn of a malachite ore from Katanga contg 0 36% Cu resulted in high cyanide consumption Pice regeneration was impossible Details are given ALDEN H EMERY

Magnetite in copper mats F G HAWLEY *Eng Mining J* 131, 319-22 (1931) — Magnetite causes Cu losses in the slag from the smelting of mat This study of the distribution of the magnetite should help in the soln of the problem and also demonstrate the value of improved chem and microscopic methods of analysis Analytical data are discussed The sulfide is sepd from the magnetite as follows Dissolve 30 g $NaClO_4$ in 50 cc hot water Cool and add 50 cc of HNO_3 Cool again and add 30 cc of 85% H_2SO_4 and cool to room temp or lower The $NaClO_4$ must be entirely in soln when the acid is added Weigh 0 5 g of mat into a 600 cc beaker and add 8-10 drops of water Spread the mat over the bottom of the beaker by shaking, place in a cooling trough and cautiously add 25 cc of the chlorate In a few min heat the beaker nearly to boiling add 50 cc of cold water and immediately filter and wash The magnetite may be weighed as such and then a correction made for the impurities, or it may be ignited and weighed as Fe_2O_3 and impurities deducted, or the Fe may be detd volumetrically and called to Fe_2O_3 W H BOYNTON

Significance of the Dwight-Lloyd method for working up of Rammelsberg ores. W SALERBERG *Metall u Erz* 28, 101-11 (1931) — The advantage of the Dwight Lloyd method for roasting Rammelsberg ores (galena, zinc blende, Cu pyrites, Fe pyrites and heavy spar) is primarily in the fact that the process is uninterrupted The capacity is also great and the efficiency good the operating cost for the roasting and sintering of a ton of raw ore being about a third less than the usual roasting and briquetting methods Another advantage is in the handling of large quantities of raw materials without a preliminary roasting, and in making available the SO_2 gas for H_2SO_4 manuf As a result of the increase in furnace capacity and the saving of coke, together with a more favorable yield of metal, the total cost has been lowered considerably H STOFRTZ

Milling methods at the Hughesville concentrator of the St. Joseph Lead Co., Hughesville, Mont. Wm O VANDERBURG *Bur Mines, Information Circ.* 6447, 15 pp (1931) — Galena, sphalerite, marmatite and pyrite occur in a fissure vein in syenite. The first 2 are argentiferous A small quantity of Cu occurs as cupriferous pyrite and chalcopyrite The gang consists of altered syenite and rhyolite with subordinate amts of calcite, barite, quartz, rhodochrosite and marcasite The av. grade of the ore is Pb 6 21, Zn 5 01% and Ag 9 09 oz/ton By selective flotation 2 concentrates are made: one Pb 61, Zn 5 7% Ag 50 oz and Au 0 05 oz per ton and another Zn 51, Pb 1%, Ag 31 oz and Au 0 02 oz/ton Pb recovery is about 97%, Zn about 92% and Ag 80% ALDEN H EMERY

Milling methods at the Hurley plant of the Nevada Consolidated Copper Co.

Hurley, New Mexico. FRED HODGES. *Bur Mines, Information Circ* 6394, 16 pp (1931). The Cu mineral is chiefly chalcocite, although this is in places accompanied by azurite malachite chrysocolla and cuprite. The gang is quartz, pyrite, sericite and halloysite. For the first 4 months of 1930 heads averaged 1.32% Cu and 3.14% Fe. Recovery was 53.9%. Concentrates assayed 27.51% Cu, 0.04 oz Au and 0.55 oz Ag per ton. Ore is concentrated by flotation with lime, rosin and pine oil. A. H. F.

Metal working in power presses. E. V. CRANE. *Am Inst Mining Met. Eng. Inst. of Metals Division* (preprint) 1931, New York meeting, 31 pp.—Metal working in presses is discussed under the subdivisions of shearing, bending, drawing, stamping, squeezing, extrusion and forging from both the metallurgical and stress analysis viewpoints. The stamping trade does its engineering largely by trial and error. The principal operations of metal forming have many features in common with rolling and wire drawing methods. The division between hot and cold working is taken as the recrystallization range of temp. of the metal being worked. The specific features of the different operations and the factors involved are analyzed in some detail. C. L. MANTELL.

Recent design of quicksilver plants. C. N. SCHUETTE. *Eng Mining J.* 131, 316-5 (1931).—Three types of furnaces are in common use—the old Scott furnace, the rotary kiln and the Herreshoff furnace. The Scott is unsuited to intermittent operation and requires a reasonably dry feed. The Herreshoff furnace has the advantage of a higher fuel economy than the rotary kiln and no occasional shutdowns for relining are necessary. Dust elimination is discussed and several illustrations from installations at different mines are shown. W. H. BOYNTON.

Reduction of zinc oxide by methane or natural gas. H. A. DOERNER. *Bur Mines, Rept. of Investigations* 3091, 14 pp (1931).— $\text{ZnO} + \text{CH}_4 \rightarrow \text{Zn (gas)} + \text{CO} + 2\text{H}_2$ is nearly complete when equal is attained (O gas velocity) at about 927°. The reaction rate increases with temp., but CO_2 formation and subsequent reoxidation of Zn also increase. Temps should therefore be kept below 1000°. Crude natural gas gives approx. the same results as C_2H_6 , the higher reducing power of the homologs of CH_4 being compensated by their greater tendency toward thermal decompos. A preliminary treatment to remove (sulfate) S is advantageous. This may be accomplished by heating the calcine at 800° in an atm. of natural gas or the gaseous products from the Zn condenser. Cd may also be eliminated by this step. In the temp. range 975-1000° the rate of Zn distn. compares favorably with that obtained from retorts by standard practice; the metal produced is of exceptional purity and no appreciable amt. of blue powder is formed. The special requirements of a furnace for concn. reduction of Zn by natural gas are (1) a retort and condensing system as nearly gas-tight as possible, (2) continuous feed and discharge of ore at a slow rate (time of residence about 34 hr.), (3) continuous counterflow of gas at a low velocity (time of contact, about 1 min.), (4) retorts having a large capacity, and (5) pretreatment of the ore in a sep. retort at 800° or lower with gas from the Zn metal condenser. Under these conditions Cd and S are volatilized and Fe is reduced to metal. The length of the retort is limited by the resistance of the ore to the flow of gas and by the necessity to prevent transportation of dust incident to the highest gas velocity required by long retorts if the time of contact is constant. The smallest dimension of the retort will be limited by the requirements of heat transfer. The relatively low operating temp. and the presence of the excellent thermal conductor, H_2 , facilitate the flow of heat and permit large units to be used. ALDEN H. EMERY.

Iron and steel in 1930. ANON. *Mining J. (London) Annual Rev. No.* (Jan. 24, 1931) 1. Copper in 1930. EDWARD H. ROBBIE. *Ibid* 2. Tin in 1930. E. BALLOL SCOTT. *Ibid* 3-4. Lead and spelter in 1930. P. W. SMITH. *Ibid* 5. Gold and silver in 1930. ANON. *Ibid* 7. Platinum in 1930. ANON. *Ibid* 8. Cadmium. W. C. REMBOLD. *Ibid* 13. Modern industrial applications of steel. ANON. *Ibid* 15, 17, 19, 21-2.—The following uses of steel are discussed: the constructional industry, high-pressure and high-temp. vessels, chem. and allied industries, automobile and aircraft construction, mine shafts and underground supports, rock drilling, railroad, bridge work, trucking and pulverizing machinery. The use of alloy metals in iron and steel. ANON. *Ibid* 23.

New views on smelting [of iron ores] and remelting processes. J. E. FLETCHER. *Foundry Trade J.* 44, 175-6, 195-7, 211-3, 227-9 (1931).—See C. A. 24, 5263. D. S.

Smelting in the lead blast furnace. Handling rich charges. VI. Conditions and problems introduced by increasing the ratio of concentration. G. L. OLDRIGHT AND VIRGIN MULLER. *Bur Mines, Rept. of Investigations* 3088, 7 pp (1931), cf. C. A. 24, 219.

Design of open-hearth ports. V. BALABANOV. *Iron Age* 127, 1254-7, 1273 (1931). E. H.

Graphical carbon balance of the blast furnace. T J ESS *Iron Steel Eng* 8, 173-6(1931) —Charts developed to facilitate a check on gas production, blast vol and furnace practice are shown and the application is explained in detg the carbon balance. The balance, worked through analytically, agrees substantially within slide rule limits.

W H BOYNTON

Carbon can be regulated by cupola practice. GOTTFRIED OLSON *Foundry* 59, 62-3 (1931) —In order to control the C in cupola low C gray irons it is necessary to have (1) a forehearth, (2) low tuyères, (3) correct tuyère ratio, (4) proper amt of air supplied under proper pressure, (5) rapid melting, (6) high temp in melting zone, (7) large amts of steel scrap, (8) coke with approx 7% ash, (9) coke with moisture under 1%.

DOWNS SCHAAF

Regularities in the composition of basic Siemens-Martin slags. SIGFRIED SCHILFICHER *Arch Eisenhüttenw* 4, 239-44(1930) —From the analyses of 29 slags it appears that the MnO content depends on the value of CaO + MgO. The sum of the 3 values is practically const at 61.5%. Furthermore, the Fe content decreases with increasing SiO₂ and P₂O₅. If the Fe and MnO contents of a slag are known the approx compn can be calcd.

H F JOHNSTONE

Application of cementation methods of the systems Sb-Cd, Zn-Cd, Pb-Cd, Sn-Cd and Bi-Cd. WL LOSKIEWICZ *Przegląd Tech* 69, 508-13(1930) —A cementation method as described elsewhere (*Przegląd Górnictwo Hutniczy* 1929, 583) was employed in investigating the cementing characteristics of the systems Sb-Cd, Zn-Cd, Pb-Cd, Sn-Cd and Bi-Cd, which on cementation below the eutectic temp form permanent openings at the junction. By cementation at a temp above the eutectic temp the eutectic is formed, this is followed by fusion at the junction, which prevents permanent openings. Tests carried out below the respective eutectic temps showed no evidence of cementation, while those cemented 3% above the eutectic temp showed fusion and quantities of the eutectic. Methods employed are discussed. In the Sb-Cd system Cd-Sb is formed at the junction, while in the other systems investigated, eutectics were formed in every case. In the system Bi-Cd, there exists a fusion in a stable form and in the other systems tests corroborate the existence of such fusion.

FRANK KOZOSKI

The constitution of the iron-carbon-silicon-system. II. Section through the tridimensional diagram at 8% silicon. A KŘÍŽ AND F POBOŘIL. *Collection Czechoslov Chem Comm* 3, 61-72(1931), cf *C A* 24, 5705 —The properties of 7 alloys were investigated. These alloys were prep'd by melting in an elec resistance furnace. The percentage chem compn. of the alloys tested is

Alloy No.	C	Mn	Si	P	S	Ac ₁ in degrees
34	0.14	0.43	7.97	0.041	0.011	645
35	0.21	0.31	7.87	0.045	0.011	660
36	0.34	0.40	7.63	0.062	0.015	645
37	0.50	0.48	8.12	0.038	0.016	650
38	0.86	0.38	7.94	0.062	0.023	650
39	1.88	0.46	7.45	0.052	0.024	665
40	2.7	0.55	7.84	0.097	0.009	665

Samples for thermal analysis and microscopic examn after heat treatment were taken from each alloy. The constitutional diagram of the ternary Fe-C-Si system contg 8% Si is given. The region of the homogeneous γ phase does not exist in these alloys. The pearlitic transformation of the alloys with more than 0.48% C ends at the beginning of melting. The transformation Ac₁ in the alloys with 8% Si lies in the region $\alpha + C$ at 665°.

G T. MOTOK

Practical use of the knowledge of structure of steels. F GONZÁLEZ V. *Rev quim* 6, No 2, 21-5(1930) —Based on the chem compn of the steel an attempt is made to calc all the structural phases existing in steel such as free pearlite, Mn₃C, MnS, Fe₃P and free ferrite. The Juptner, Neil and Osmond formulas have been used in calcg tensile strength, hardness and elasticity.

G T. MOTOK

Corrosion and metal protection in steam power plants. IV. R. STUMPER. *Korrosion Metallschutz* 7, 25-8(1931), cf *C A* 25, 1206 —A boiler-tube failure from a gas fired furnace using a hard water was investigated metallographically. A deposit of CaSO₄ was found to exist throughout the entire length of the tube. Sections of the tube were cut from the tube at the point of failure and at a point some distance from the blow out. Microscopic examns of the metal surface of both sections revealed decarburization of the outer side of the section which failed, and indicated that the temp of the inside of the tube had been between the Ac₁ and Ac₃ points, while temp conditions in the other portions of the tube had apparently been such as to cause no changes in the micro-

structure. The conclusion was reached that localized overheating due to an impingement flame had caused the failure. B I ROPTIN 11

The corrosion of iron by steam at high temperatures and its physico-chemical basis. I W KRAUS *Korrosion Metallschutz* 7, 29-34 (1971).—A method is described for measuring the rate of oxidation of Fe in superheated steam at high times. The method consists in measuring the increase in elec. resistance of a metal wire during the process of oxidation. The resistance $R = \frac{\rho p \pi l}{\pi} \left[\frac{1}{\rho_1 r_1^2 + \rho_2 (r_1^2 - r_2^2)} \right]$, in which

r and r_2 are radii of the bare wire and the oxide covered wire, ρ_1 and ρ_2 the sp. resistances of the metal and the oxide, ρ is the ratio of molal vols. of the metal and the oxide, and l the length of the wire. Measurements of resistance changes made at temps. ranging from 400 to 800° and detns. of O_2 in the oxide film showed the ratio of O_2 used to Fe reacting to be approx. that found in Fe $_2O_3$. Sp. detns. of wt. increases due to oxidation of Fe in furnace at the same temps. were found to agree well with those calculated from the resistance changes. Both methods indicated that the thickness of the oxide layer increased with time according to the diffusion equation $x^2 = kt$ in which x = thickness in μ , t the time in min. The values of k were detd. for temps. of 500°, 700°, 800° and 1000° and were resp. 8.08×10^{-4} , 1.77×10^{-4} , 1.22×10^{-4} and 1.27×10^{-5} . Temp. coeffs. for the reaction rate were found to be 3% per degree between 500° and 700° and 1.5% per degree between 700° and 800°. B I ROPTIN 11

Corrosion of iron by carbon tetrachloride. J MILHAUER *Collection Czechoslov. Chem. Comm.* 3, 73-75 (1931). Iron is corroded by CCl_4 in presence of EtOH and water. A test made consisted in boiling iron for 1 hr. in a mixt. of CCl_4 , EtOH and water. The results obtained were

CCl_4	EtOH	H $_2$ O	% Fe dissolved
80	11	9	0.21
50	41	9	0.33
30	60	10	0.0
20	18	2	1.04
15	5	50	0.13
10	9.1	0.0	0.63
70	27.3	2.7	1.81
70	20	10	0.74
70	10	20	0.10
100	0	0	0
0	100	0	0

Max. corrosion is obtained with a mixt. of 70% CCl_4 , 27.3% EtOH and 2.7% water. The presence of EtOH favors the corrosion of Fe by CCl_4 . G T, MOTOR

Problems of insulation [of metals against access of moisture] (MOLL) 13. Some factors of solidification in relation to metallurgical problems (SMITH) 8. New methods of coke testing (MILLER) 21. The molecular constitution of the α solid solutions of Sn in Cu (JEFFERY) 2. The molecular constitution of the β solid solutions of Sn in Cu examined thermodynamically (JEFFERY) 2. Detonators [allows for their synthesis] (Fr. pat. 696,663) 24. Leaching materials (Ger. pat. 517,930) 13. Rotary fuel furnace and cooler for ores (Ger. pat. 520,212) 20. Apparatus for cleaning blast furnace gases (U. S. pat. 1,797,908) 1.

BARTELS, W. Die Dauerfestigkeit ungeschweisster und geschweisster Guss- und Walzwerkstoffe. Berlin V D I Buchhandlung M 3.

BECKER, ERICH Die Formpraxis in der Metallgießerei. Halle W. Knapp 114 pp. M 7, bound M 8.50.

BECKER, ERICH Die Gieß- und Putztechnik in der Metallgießerei. Halle W. Knapp 68 pp. M 4.20, bound M 5.50.

Bericht über die Versorgung der deutschen Wirtschaft mit Nichteisenmetallen Through June, 1930. Berlin V D I Buchhandlung Bound M 4.90.

GUILLET, L. Trempe, recuit, revenu. T. III. Résultats. Paris Dunod 490 pp. F 170, bound F 180. Cf. C. A. 23, 365.

HORNBERG, JOSEPH G. Practical Iron Founding. Revised by Walter J. May. 5th ed. London Sir Isaac Pitman and Sons Ltd. 409 pp. 10s. net.

MARCUS, ALFRED Die grossen Eisen- und Metallkonzerne. Leipzig Verlag S. Hirzel 192 pp. M 13, half linen, M 14.50.

METALLUR. CHRISTIAN A. Untersuchungen über die Seligerung der Begleitelemente des Eisens (thesis) Technische Hochschule, Aachen, Germany 25 pp. Reviewed in *Metals and Alloys* 1, 632 (1930)

NICKEL, J. S. Literature and Patent References to Platinum Group Metals. Vol. I, Nos 1, 2 and 3 1p 1 21, 22-58, 69-105 New York International Nickel Co

SALAN W. Die Dauerfestigkeit der Leichtmetall-Sandguss-Legierungen. Berlin V D I Buchhandlung M 665

SCHLESINGER A. Die ungünstige Lage der Fertigwarenindustrie in Nicht-Eisenmetallen und ihre Gründe. Berlin V D I Buchhandlung M 8

Testing and Grading Foundry Sands. Chicago Am Foundrymen's Assoc \$3
Reviewed in *Foundry Trade J* 44, 270 (1931)

Aerating apparatus for flotation plant for ores, etc. CENTRAL-EUROPAISCHE SCHWIMM-APPARATE F. A. G. (R C Gibbs, inventor) Ger. 519,033, Feb 2, 1929
Addn to 1,111 (C A 25, 1209)

Concentrating ores. RICHARD B. MARTIN U S 1,707,355, March 24. In effecting flotation of such as Florida phosphate ore screenings with selective agents such as oil and Na₂CO₃ the water content of the ore is regulated to the amt necessary to wet all the ore particles without having present an excess which will settle out on the tank and the wetted ore is tumbled with the selective agents to prepare it for flotation.

Electromagnetic separator for ores, etc. WM M MORRIS Ger 521,020, May 10, 1924. Corr spoils to Brit 221,921

Briquetting ores. FRIEDRICH ALBERT KNAPP and LEON MAYER Ger 517,735 April 1, 1924. Addn to 501,800 (C A 24, 4752) The method of 501,585 for briquetting materials with C and low stage oxidation compds of Fe, is extended to ores of Cr, W, V, Mn, Ti, Ni, Co and Pb

Treating ores, etc. ERICH KATZ GROSSWITZ A G Ger 519,484, April 13, 1926. Addn to 473,016 (C A 23, 2899) The method of Ger 473,016 is modified by supplying the materials to the furnace in a pre-warmed state, or by heating the end of the furnace at which the materials enter. The reaction zone is thus extended toward that end of the furnace and this facilitates the completion of the volatilization before the residue in the furnace softens or sinters.

Treating sulfide ores, etc. HAROLD SHAPIRO U S 1,707,700, March 24. Material such as sulfided heavy metal ore is subjected to a reducing smelting operation in the presence of light metal compds such as those of Ca or Mg, in such quantity and with such regulated conditions of reduction as to produce a slag having an Fe content of below 15% FeO and a max contg light metals. Numerous details and examples are given.

Sulfide ores. SULPHIDE CORP. Fr 696,520, May 22, 1930. Ores contg Fe and S are treated for the production of S, oxide of Fe and chlorides of other metals contained in the ores by forming FeCl₃ with displacement of S in the form of vapors, followed by oxidation of the FeCl₃ to Fe₂O₃ and FeCl₂, the latter distd with N being used as a chlorinating agent for displacing S. The temp in the chlorinating and oxidizing gases is maintained above the combination pt of the S or of the FeCl₃, but below the temp which would produce fusion of the solids present.

Chromium ores. I G FARBWIND A G (Paul Weise and Julius Drucker, inventors) Ger 511,743, Apr 17, 1927. Cr ores which have been heated and quenched are decomposed by subjecting them to the action of gases contg HCl, Cl and reducing agents may be present in the gas. Thus, Cr ore is heated to 800° and quenched in water. The quenched ore is subjected to the action of gases contg HCl and a reducing agent at 550-600°, leaving moderately pure CrCl₃. Other examples are given.

Extraction of iron from its ores. VEREINIGTE STAHLWERKE A-G Ger. 517,000, June 8, 1927. The ores are treated with Cl-HCl gaseous chlorides, etc., and the resulting FeCl₃ is decomposed to obtain the Fe. The FeCl₃ is heated to above its dissociation temp in an elec arc and the decomposn products, Fe and Cl, are cooled and sep'd.

Extracting copper. ANGLO-AMERICAN CORP OF SOUTH AFRICA, LTD Ger. 521,184, May 19, 1928. See Brit 501,850 (C A 23, 4181)

Extracting copper, etc. PATENTAKTIEBOLAGET GRÖNDAL-RAMEN Ger 521,111, July 10, 1929. Ores contg Cu and Zn are treated with SO₂ after addn of chlorides if necessary, until the HCl content is 5-10 g per l. Part of the Cu is thus pptd as CuCl, and the pptn is then completed with Cu. The lye and the ppt are then treated with Fe,

either separately or together to recover Cu. The method may be applied to the lyes obtained by extg. pyrites that have been roasted under chloridizing conditions, and if Ag and Au are present they may be pptd. with the CuCl by using an excess of Cu.

Nickel and copper. THE INTERNATIONAL NICKEL CO., INC. Fr 696,456, May 8, 1930. Ni and Cu are extd. from materials contg. them, particularly mats contg. sulfides of Ni and Cu by heating the materials in the presence of a flux to obtain a molten mass and stratifying to sep. the material contg. Ni from that contg. Cu. The part contg. Cu is heated in the presence of air to obtain porous Cu and the part contg. Ni is submitted to the carbonyl process to obtain Ni.

Reduction of tin. BEZELIUS' METALLURGIEN G. M. B. H. Fr 696,146, May 27, 1930. Finely granulated materials contg. SnO_2 are reduced by solid, liquid or gaseous agents at a temp. kept below the m. p. of the charge during the reduction and raised at the end of the reduction sufficiently to give a molten or pasty slag which surrounds the reduced Sn and preserves it from reoxidation in the furnace itself or during a subsequent treatment.

Extracting zinc, etc. HANNS BEERMANN. Ger 521,485, Nov 7, 1929. In extg. Zn and other volatile metals from a mixt. of oxide ore and C in a horizontal, externally heated muffle a porous or perforated tube, to which gas is supplied, is arranged along the whole length of the bottom of the muffle. The charge is loosened and agitated by the fine streams of gas ascending through it.

Zinc. THE NEW JERSEY ZINC CO. Fr 697,217, June 11, 1930. Cd and the like are removed from zinciferous materials by heating the materials in the presence of carefully regulated amts. of a chlorinating agent, steam and air and the treatment is selectively regulated to volatilize the Cd without appreciable loss of Zn. Cf. C. A., 25, 1927.

Roasting pyrites, etc. ALEXANDRE FOLLIET and NICOLAS SAINDERICHY. Fr 697,006, May 26, 1930. Pyrites and S contg. ferruginous materials are roasted and agglomerated at the same time, by the progressive action of air, heated in a special app. to temps. of 650-800°, on the material advancing progressively in a thin layer in a rotating or other furnace.

Treating roasted pyrites. REYMERSHOLMS GAMLA INDUSTRI AKTIEBOLAG. Ger 517,748, Nov 13, 1925. The valuable constituents of chlorinated roasted pyrites are obtained by leaching the pyrites in stages with (1) lye poor in Cu but rich in Zn, contg. mainly chlorides and (2) lye poor in Zn but rich in Cu, contg. mainly sulfates. Cu, Zn, Ag, Co, S and Cl compds. are so obtained.

Rotary ore-roasting and sintering apparatus. GEORGE E. STORER (to National Processes, Ltd.). U. S. 1,797,931, March 24. Structural features.

Refining metals. SYNDICAT DES LAITIERS ET SCORIES. Fr 696,770, Sept. 11, 1929. Loss of metals in slags during the refining of ferrous metals is reduced by adding to the slag, during or after refining suitable compds. such as oxides, carbonates, chlorides, borates or sulfates of alkali or alk. earth metals or Mg.

Refining metals. EMIL VIDAL and ARNOLD H. STROCKHAUSEN. Fr 695,623, May 13, 1930. A system of extg. and refining metals is described in which only schist heated by oil burners fed by compressed air liberates oxidizing and reducing gases to the ore which, in the form of briquets or broken, is placed above the schist in a refractory brick furnace and covered with a layer of lime. The metal in fusion falls through a grid of refractory bricks to wells which are heated from below. The combustion products are delivered to clarification, condensation and extra chambers.

Volatilizable metals. METALLURGES A. G. Fr 696,151, May 27, 1930. Metals such as Pb and Zn are volatilized by introducing the metal contg. material mixed with a reducing substance into a rotating furnace and causing it to move continuously in the furnace. The air is introduced into the reaction chamber in amt. sufficiently low to insure a reducing atm. in the reaction chamber. The material is preliminarily heated by the complementary combustion of the gases from the reaction chamber.

Protecting metals. C. F. BOEHRINGER & SOHN G. M. B. H. Fr 695,787, May 17, 1930. Metals and their alloys are protected against destructive action by adding quinodine or substances contg. quinodine to the agents coming in contact with the metals.

Porous masses of metal or metal oxides. I. G. FARRINGTON A. G. Brit 339,645, Sept. 5, 1929. A foam is formed from a mixt. of a liquid, a foam producing agent and finely divided metals or metal oxides (or compds. which are converted into metal oxides when heated) and the foam is dried. Numerous details and examples are given, and according to the materials used the products may be suitable for use as bearing metals, storage battery plates or fillings for them, filters for liquids or gases, fillers for building

stones or bricks, diaphragms, molds for casting gypsum, glass or metals, backfire preventing devices for gas burners, fillings for teeth, etc

Metal founding. FRITZ SINGER Ger 516,575, July 23, 1926 A protecting layer is formed on molds, especially for Al, by coating with Zn or Mg and heating The Zn or Mg is applied as dust, optionally with a binding agent

Foundry molds and cores JFAN B DURAND Ger 520,175, April 5, 1929. A mortar is used which consists of a hydraulic binder, sand and an amt of water insufficient for the complete hydration of the binder A suitable compn is fine portland cement 100 sand 600 and water 90 pts by vol Cf C A 24, 2416

Core-inserting or -removing apparatus for molds for centrifugal casting. CENTRIFUGAL CASTINGS LTD Ger 519,236, Dec 11 1928 Corresponds to Brit 301,015

Double-walled ingot mold. HUBERT KAMPS and HERMANN ZEPERNICK Ger 519 235 April 9, 1929

Water-cooled mold for casting roller plates ANTON MOHR Ger 517,780, Feb 1, 1930

Molds for casting iron and steel WILHELM KLEFSCH Ger 520,232, June 30, 1926 See U S 1,777,975 (C A 24, 5710)

Block mold for casting steel. MAXWELL G DEMAS Ger 517,778, June 2, 1928

Apparatus for casting metals such as aluminum. BIRMINGHAM ALUMINUM CASTING (1904) CO, LTD, and C VAUGHAN Brit 339,721, May 22, 1930 Structural features

Casting aluminum and its alloys. H RÖHRIG Brit 339,624, July 8, 1929 For obtaining a uniform crystal structure in an ingot of Al or Al alloy, the mold is cooled at the bottom only, the top and sides being heated or packed with heat insulating material

Apparatus for centrifugal casting of metal pipes, etc. J A BYERS and M C SMOTZER (to Centrifugal Pipe Corp) Brit 340,137, Dec 31, 1928 Structural features and various details of operation are described

Centrifugal casting of metal pipes. LEON L MATHIEU (to Soc anon des hauts fourneaux et fonderies de Pont à Mousson) U S 1,707,872, March 24 Mech features

Centrifugal casting of metal pipes, etc. A POSSENTI and C SCORZA Brit 340, 129, April 12, 1929 App and various mech features are described

Method of making hollow bodies by centrifugal casting. MANNESMANNRÖHRENWERKE Ger 519 237, June 16, 1928

Casting carbides of metals difficult to fuse. GEWERKSCHAFT WALLRAM Ger 516,650, Sept 9, 1924 Metals and metalloids difficult to fuse, such as W, B, etc, are superheated in a crucible with C, with such rapidity that an excessive amt of C is not taken up by the fused metal in forming the carbide The carbide is cast in the usual way Cf C A 24, 5211

Furnace for fritting powdered ores and metallurgical products. VEREINIGTE STAHLWERKE A-G Ir 695,502, June 2, 1930

Lining metallurgical furnaces. HIRSCH, KUPFER- UND MESSINGWERKE A-G Ger 521,295, June 6, 1928 An elec conductive core corresponding in shape to the furnace interior and having a higher m p than the lining compn is introduced into the furnace The lining compn is placed in the space between the furnace wall and the core, and the latter is then electrically heated, directly or inductively, to a temp above the m p or sintering temp of the lining

Cupola and like furnaces. JOSEPH E FLETCHER and THE BRITISH CAST IRON RESEARCH ASSOCIATION Ir 695,555, May 14, 1930 Means for controlling admission of air to the cupola is described

Regenerative furnaces for fusing metals. WALTER ALBERTS and PAUL ZIMMERMANN Ir 695,830, May 19, 1930

Method and apparatus for reducing dust losses in blast-furnace operation. GEORG EICHENBERG and NIKOLAUS WARK Ger 520,164, Sept 15, 1929

Water-cooling system for blast furnaces, etc. B J MULLEN Brit 339,972, Dec 19, 1928 A signal system (details of which are described) is provided for operation when leakage of gas into the water cooling system causes fall in the liquid level

Furnace for reclaiming scrap metal. THERON D STAY and HARRISON O BURROWS (to Aluminum Co of Am) U S 1,797,276, March 24

Furnace for heating metal sheets and pairs. CHARLES A DAVIS and FRANK I. LEAHY. U S 1,797,902, March 24 Structural features

Cast iron. MEIER & WEICHELDT. (Emil Schurz, inventor). Ger. 521,432, May 14, 1924 Gray cast iron with a tensile strength above 28 kg per sq mm is prep'd by casting com. Fe contg about 3.5% of Si under such conditions of regulated cooling that

the free C is uniformly distributed in the metal as graphite Fe eutectic. Casting may be effected in Fe molds.

Spongy iron. IVAR RENNERFELT and IVAR C S G BOHM. *Ir* 697,078, June 3, 1930. Spongy Fe is made by introducing a mixt of Fe ore and a carboniferous material into a furnace similar to a coke oven heated on the exterior by gas. Toward the end of the operation gas is introduced into the furnace itself to absorb the remaining O. The cake of spongy Fe is removed and introduced into a case kept free from air. The reduction may be completed by the gas introduced.

Iron-reinforced converter floor for steel manufacture. MANNESMANNROHRENWERKE. *Ger* 517,814, May 8, 1929.

Steel. FREDERICK F MCINTOSH to Crucible Steel Co of Am. *U S* 1,797,728, 24. In melting and refining steel the product is withdrawn from the refining lad and while still molten there is added to it a predetd quantity of Mo sulfide, in order to effect combination of both S and Mo with the steel.

Phosphorus steel. F BURGGRAB. *Brit* 339,016, Dec 22, 1928. Easily machinable steel for pressed nuts etc., of high P content, is prepd by first melting down in a cupola a mixt of scrap iron, siliceous pig iron and a material rich in P such as basic pig iron, apatite or vivianite or ferro P (with adjustment of the P content according to the product desired) and decarburizing the molten in an acid lined Siemens Martin furnace. Ferro-Mn may be added as a decarburizing agent. The steel produced may contain P 0.2-0.5, Mn 0.3-0.6, C 0.04-0.1, S 0.005-0.15 and Si up to 0.25%.

Rustless steel. VEREINIGTE STAHLWERKE A G. *Ir* 695,395, May 10, 1930. A steel or cast iron which does not tend to rust is made by alloying, with the steel or Fe, elements whose electrolytic potential is below the const. oxidation potential of the corrosion agent but is above the electrolytic potential of the steel or Fe, and which by reason of their electrochem properties provoke during the corrosion, the formation of an adhering anti rusting layer. Sb, As, Bi or Sn in amts of 0.5-5% alone or together and with or without Cu may be used.

Apparatus for heat-treating steel rolls for rolling mills. CHARLES F LOGAN (to Union Elec Steel Corp). *U S* 1,797,430, March 24. Structural features.

Alloys. COMPAGNIE FRANÇAISE POUR L'EXPLOITATION DES PROCÉDÉS THOMSON-HOUSTON. *Fr* 606,142, May 27, 1930. Alloys are made by intimately mixing easily reducible compds of the constituent metals, and reducing the mixt to the metallic state.

Alloy suitable for jewelry. HARRY KLAUSMANN and HENRY R KIERN (to Baker & Co). *U S* 1,797,236, March 24. An alloy, which is whiter and harder than Pd, is formed from Pd 75-88 and Rh 21-15%, the remainder being Ru (which may be present in the proportion of 2%).

Aluminum alloys. THEODORE W BUSNERT and JOSEPH A NOCK, JR (to Aluminum Co of Am). *U S* 1,797,861, March 24. Alloys which are suitable for mech working comprise Al at least 95% together with Mn 1-1.5 and Mg 0.5-1.5%, free from other substances except such as are present as incidental impurities.

Aluminum alloy. OTTO KAMPE. *Ir* 695,838, May 19, 1930. An alloy of Al contains Ag, Cu and Ti. V, Pd and Ni may also be added. An example contains Cu 2-5 or Ni 0.2-0.8 or Cu and Ni together 2-5, Ag 0.2-1, Ti 0.02-0.2, V or Pd about 0.1% and Al the rest. The stability on annealing of the alloy is increased by submitting the alloy in a liquid state and at a temp of about 750° to a refining with SrCl₂. The SrCl₂ may be introduced to the fused mass in an Al capsule. *Cl C A* 25,2410.

Improving iron-chromium-nickel-carbon alloys. FRIED KRUPP A-G. *Ir* 697,188, June 11, 1930. The resistance values of austenitic alloys of Fe-Cr-Ni-C (e.g. Cr 6-40, Ni 40-4 and C 1%) are improved by submitting these alloys to a sudden tempering from a temp below 1000°.

Iron-nickel alloy for vapor burners. AKTIEBOLAGET B A HJORTH & Co. *Brit* 339,831, Feb 4, 1929. Double-walled caps forming the mixing chambers of vapor burners are made at least in part of an alloy comprising Fe 62-67 and Ni 38-33% which is stated to have a suitably low heat cond.

Ferrosilicon. I G FARSENAUD A G (Carl Müller, Alfred Curs and Leo Schlecht, inventors). *Ger* 519,066, Apr 9, 1926. Fe-Si alloys free from C or of predetd C content are prepd from Fe obtained by decomp Fe(CO)₅.

Magnesium alloys for pistons. I G FARSENAUD A G. *Ger* 519,377, Dec 6, 1928. Alloys for making pistons of internal combustion engines comprise Zn 1.8-5 and Si up to 6, with or without Al, Cu and (or) Sn up to a combined total of 5%, the remainder being Mg. *Cl C A* 25,2109.

Manganese steels. TAYLOR WHARTON IRON & STEEL CO. *Fr* 695,871, May 20, 1930. See *Brit* 336,091 (*C A* 25,1792).

Steel alloys. FRIED KRUPP A-G and HEINRICH KOPPERS A-G Fr. 697,053, June 2, 1930 Steel alloys resistant to chem. agents, *e.g.* solns. contg. H_2SO_4 contain Cr 17-40, Ni 6-25 and C up to 0.5%, and particularly Cr 17-18, Ni 8-10 and C 0.02-0.2%, with or without Mo 0.2-3%.

Zinc alloys. E A ANDERSON (to New Jersey Zinc Co) Brit 340,104, June 5, 1929 Alloys suitable for die-casting are formed of Zn (of about 99.99% purity) together with an addn. of Al 2-10 and Mg 0.01-0.3% (the Pb and Cd together not being more than about 0.01%, preferably not more than 0.003% each of Cd and Pb and not more than 0.001% Sn). Zn alloys are also mentioned contg. Al 4 and Cu 3%, with or without Mg 0.1%. Cf C A 25, 66.

Cleaning metallic surfaces. JAMES H GRAVELL Fr. 695,914, May 21, 1930 Metallic surfaces are exposed to the action of a soln. which contains a compd. of H_3PO_4 , one atom of H of which is replaced by a metal or radical such as NaH_2PO_4 or $\text{NH}_4\text{H}_2\text{PO}_4$.

Rustproofing metals. CHEMIEPRODUKTE G M B H Ger. 517,579, Oct. 18, 1928 Metal parts are protected against rusting when under water by giving them a ground coating of bitumen etc. to which is then added a difficultly sol. inorg. substance which will readily enter into ionic exchange with bases. Thus, the parts are coated with bitumen to which is added a coating of HgSiO_3 . Further examples are given.

Rust-proofing iron, steel, etc. WM H COLE Ger. 516,729, Mar. 19, 1928 See Brit 292,666, (C A 23, 1332).

Coating metals with refractory materials. DOHERTY RESEARCH CO Fr. 696,422, Apr. 8, 1930 Metals are coated by applying a refractory agglomerate to the surface and the free surface of this agglomerate is exposed to a high temp. to calcine or harden it, the subjacent parts of the agglomerate being exposed to progressively decreasing temps., and the metal is protected against exposure to a high temp. A binding substance capable of attacking the metal when heated may be added to the refractory material so that the products of the reaction, *e.g.* the oxide of the metal, penetrate the refractory material which may be a compd. of Zr.

Apparatus for annealing metal objects. AKT-GES. BROWN, BOVERI & CIE SWISS 143,457, Jan. 4, 1930

Aluminum heat-transfer surfaces. VEREINIGTE ALUMINIUM WERKE A-G Brit. 339,722, Oct. 11, 1929 The heat absorbing or radiating capacity of articles made of Al or its alloys is increased by producing an oxidized layer on the material, *e.g.*, an engine cylinder of Mg Al alloy may be coated with a layer of pure Al and the surface of the latter then oxidized.

Nickel products such as eyeglass frames. WM J WRIGHTON and EDGAR D TILLYER (to American Optical Co) U S 1,797,254, March 24 Ni products such as eyeglass frames are shaped to the desired form by stretching beyond the elastic limit of the material, and the surface fractures and pits are then filled with Ni which has not been worked beyond its elastic limit (suitably by electrodeposition).

Non-corrosive water and alcohol solution. WM S. CALCOTT (to E I du Pont de Nemours & Co) U S 1,797,401, March 24 Corrosion of Fe and Cu by aq. alc. solns. is inhibited by adding a sol. amino carboxylic acid deriv. such as an ethanol amine salt of phenylglycine and a substance such as a mercaptothiazole such as is used as a collecting agent in the flotation of metals. U S 1,797,402 specifies a similar mixt. in which the flotation collecting agent used is an org. tervalent N compd. such as azobenzene, aminoazotoluene or aminoazotolueneazomaphthol. Numerous examples are given. Cf C A 24, 2755.

Arc-welding electrodes. J M WEED (to British Thomson Houston Co., Ltd) Brit. 339,930, July 8, 1929 An electrode is coated with an alkali metal acetate such as NaOAc in order to obtain a short arc with a low rate of fusion and deep penetrating power so as to produce a ductile weld, and the acetate may be mixed with carbonates of alk. earths such as those of Ca and Ba and with TiO_2 , and, prior to coating the electrode with flux, the electrode may be pickled with a 10% HF soln. followed by rinsing and drying and it may be preliminarily degreased with a caustic soln.

Fusible electrodes for welding. LA SOUDURE ÉLECTRIQUE AUTOGENE (SOC. ANON.) Fr. 696,592, June 3, 1930 Electrodes for welding Al and its alloys in the elec. arc are composed of an Al-Si alloy, the content of Si being 3-13%.

Welding magnesium. I G FARBERND A-G Brit. 339,923, June 7, 1929 Welding is effected with the use of a flux as described in Brit. 313,487 (C. A. 24, 1075).

Solder. SIEMENS-SCHUCKERTWERKE A-G. Ger. 520,336, Sept. 23, 1927. An approx. eutectic Ag-Pb alloy contg. Ag 0.25-5 and Pb 99.75-95% is alloyed with 0.25-10% of Cu. The solder thus obtained m. between 300° and 325°.

10—ORGANIC CHEMISTRY

CHAS. A. ROCHILLER AND CLARENCE J. WEST

Relativity and organic chemistry. W. A. WATERS. *Science Progress* 25, 627-32 (1931).—Theories concerning the structure of org. compds. are reviewed. "Only the resulting symmetry of our chem. mol's, inspected in the aggregate, is observable, and of the actual relative positions of the various atoms, one can be certain of very little."

JOSEPH S. HERNURN

The progress of organic chemistry 1924-1928. IV. Heterocyclic series. ERICH LEIDMANN. *Z. anorg. Chem.* 43, 1061-70, 1092-7, 1112-3 (1930); cf. C. A. 24, 4755. V. Natural substances of less well-known constitution. *Ibid.* 1113-0. VI. Organic compounds. *Ibid.* 1138-42. VII. New working methods of organic chemistry. *Ibid.* 1142-5. E. H.

Progress in the knowledge of oxidation processes. K. W. ROSENTHAL. *Arch. Pharm.* 269, 126-38 (1931).—This is a review on the mechanism of oxidation, principally of organic compounds. W. O. E.

Acetic acid and cellulose acetate in the United States. EVERETT P. PARTRIDGE. *Ind. Eng. Chem.* 23, 492-93 (1931).—A survey giving a general view of the AcOH industry in the U. S. and of the cellulose acetate industry as it affects the consumption of AcOH and Ac₂O. Five divisions were made as follows: (1) the general developments affecting AcOH manuf. since 1914, (2) the general technology of the various processes used or proposed for AcOH manuf., (3) the production of AcOH and its consumption in various industries during 1929, (4) the general development of the cellulose acetate as it affects the consumption of AcOH and Ac₂O. A complete bibliography of 44 references is appended and several illustrations, charts and tables are included.

W. H. BOYNTON

Partial oxidation of methane in the presence of oxides of nitrogen. DAVID F. SMITH AND RUTH T. MILLER. *Ind. Eng. Chem.* 23, 357-60 (1931).—The partial oxidation of CH₄ is markedly affected by addn. of N oxides (NO₂, NO + O₂) and HNO₃ vapor. Complete data on this reaction when carried out in glass or quartz in a small app. are reported for varying conditions of heating time (0.17-0.68 sec.), temp. (500-700°), gas compn. and amt. of catalyst. A smaller vol. of CH₄O was always obtained than that of the N oxide used, and as this latter was always destroyed if efficient conversion was had, it is concluded that the process offers no commercial possibilities for CH₄O synthesis. The mechanism is probably ionic combination of CH₄ and N oxide with subsequent decompn. to CH₄O and N₂.

ALBERT THOMAS FELLOWS

The formation of ethylene and acetylene from methane in electric discharges. KURT PETERS AND O. H. WAGNER. *Z. physik. Chem., Abt. A*, 153, 161-86 (1931).—An investigation of the optimum conditions of gas flow and compn., pressure and c. d. for the formation of C₂H₄ and C₂H₂ from CH₄ have been investigated. The min. energy consumption was 18 kw. hr. per cu. m. of C₂H₄ and 53 kw. hr. per cu. m. of C₂H₂. Under the condition of min. energy consumption, the exit gas from the discharge tube contained 4.2% C₂H₄ and 0.8% C₂H₂, whereas the max. concns. of these gases obtainable were 20 and 3% resp. The best gas mixt. for the prepn. of C₂H₄ and C₂H₂ contained CH₄ and H₂ in a 2:1 ratio. A spectral analysis of the products showed that the reaction region might be either one of 2 types. In one of these only the CH bands were visible, whereas in the other with a high energy input there were in addn. to the CH bands, the Swan spectrum, the Balmer series, and the C line spectrum. The spectral measurements were supplemented by temp. measurements in the discharge. For the reaction mechanism in the region of low energy input the known scheme for the decompn. of CH₄ into a series of H poor radicals was confirmed and the formation of the separate reaction products discussed. The presence of higher satd. and unsatd. hydrocarbons in addn. to C₂H₄ is assumed to indicate polymerization reactions resulting from the impact of 2 or more radicals. In the region in which the Swan spectrum is visible, there are 2 possibilities for the formation of C₂H₂: either the combination of 2 CH

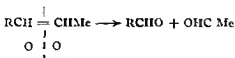
radicals, or of some reactive form of H with the $\text{—C}\equiv\text{C—}$ radical.

P. H. E.

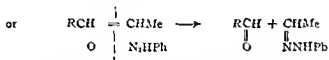
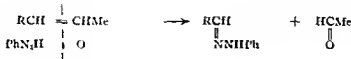
Catalytic condensation of amylenes. A. MAILHE AND RENAUDIE. *Compt. rend.* 192, 561-3 (1931).—Isomylene, Me₂CHCH=CH₂ (I), and 2-pentene (II) were each passed over silica gel in a quartz tube heated to 670° and at ordinary pressure. I yielded a gas contg. 47% ethylenic hydrocarbons (C₂H₄, C₂H₂), 38% CH₄, and 7% H₂; an oil and tar remained which on fractionation gave higher ethylenic and aromatic hydrocarbons and a residue in 80°. The gas from II was 52% C₂H₄ and C₂H₂, 34%

CH_4 and 6% H_2 , the oil and tar yielded higher ethylene and aromatic hydrocarbons and a residue m 53°. The d n and I nos of the fractions are given D S S

The oxidation of propenyl derivatives with diazo compounds. A Quilico *Atti accad Lincei* 12, 341-4(1930).—The oxidation of propenyl compds with normal diazo hydrate noted by Q (cf Q and I reri, *C A* 23, 597) was explained by the interpretation by Angeli and Polverini (*C f* 24, 4770). A shows that the general reaction is.



Thus using PhN_2OH as oxidizing agent the reaction takes place in either of the following ways



s e, forming either the alkyl or arylphenylhydrazone. The former is not isolated as it combines with unchanged azo compd forming a formazylic deriv, $\text{PhN}_2\text{CMe-N-NHPh}$

A W CONTIERI

Cuprous chloride as a catalyst for acetylene. J A NIEUWLAND AND W. L. FOONEY *Proc Indiana Acad Sci* 38, 196(1929).— C_2H_2 when run into a soln contg CuCl and NH_4Cl and satd with HCl forms Cl_2 , CHCl . C_2H_2 run into a soln contg both CuCl and CuCl_2 forms equal mixts of *cis* and *trans* dichloroethylenes. Solns of CuCl_2 in the absence of CuCl do not absorb C_2H_2 .

H M STARK

Oleostene (hydrocarbon contained in the fruit of olives). GIOVANNI SANI *Atti accad Lincei* 12, 238-42(1930).—During the extn of olive oil with alkali to remove phytosterol, an orange oil seps which is unaffected by the alkali. Etd with ether, crystals separate out, leaving the oil itself, which is insol in water and dil EtOH, but sol in other org solvents. Its EtOH soln is fluorescent d₁₇ 0.8934, analysis indicates the formula $\text{C}_{27}\text{H}_{42}$. It absorbs a max of 33.15% O_2 , indicating that the mol formula is $\text{C}_{27}\text{H}_{42}\text{O}_2$, and the mol wt of the hydrocarbon is 288, $\text{C}_{27}\text{H}_{42}$. Two N_2 derivs were also prepd which will be described in a later paper.

A W C.

The synthetic product cuprene (carbene). WALTER HERZOG *Kunststoffe* 21, 49-53(1931).—Cuprene, a polymerization product of C_2H_2 , obtained when the latter is brought in contact with Cu or CuO at elevated temps, was first prepd by Erdmann and Kothner (*Z anorg Chem* 18, 49, 57(1893)). Later investigations of this substance were carried out by Gooch and Baldwin (*Z anorg Chem* 22, 255(1900)), Alexander (*Ber* 32, 2381(1899)), Sabatier and Senderens (*Compt rend* 130, 250(1900)) and Kaufman and Schneider (*C A* 13, 1825, 16, 2302, 18, 1465). Although cuprene now is made on a large scale the technical details of the mfg process have so far remained a secret. It is known only that Cu or cupiferous bronze powders are used as catalyzers at a temp of 200-260° under pressure with the combined action of dark elec discharges. The presence of O (air) is a necessary factor. A special arrangement to prevent the voluminous polymerization products from clogging up the reaction chamber is described in Swiss Pat 95,237, May 7, 1921. Addn of 5-15% N to the C_2H_2 has been found to give homogeneous and lighter colored products. With undil C_2H_2 there is, as a result of the high reaction temp, a tendency to coking, resulting in dark products of less H content and a smaller yield. A greenish blue, neutral, strongly unsatd, tar like liquid, the "cuprene tar," d 0.89 is obtained as a by product in amts of 5.5-7% besides 13-17% waste gases. This tar has been investigated by Schlöpfer and Stadler (*C A* 20, 1384). Under the proper conditions the yield of cuprene is 25% of the wt of CaC_2 used. According to Ger pat 407,485, Dec 30, 1922, C_2H_2 , mixed with indifferent gas, is conducted into stearic acid heated to 250-300° to which has been added Cu bronze. The cuprene obtained by this process is finely powdered and of a light to dark brown color. After sepn from the stearic acid the last traces of

the latter are removed by washing with caustic lye or solvents. Addn. of air or O to C_4H_2 or addn. of substances giving off O to the contact medium, has been found advantageous. Sulfurized cuprene, contg 11% S, has been prepd. by heating cuprene with S_2Cl_2 in the presence of benzene as suspending agent under reflux. When working under 10 atm pressure and at 180-200° a product contg 13.6% S and 4% Cl was obtained. These substances may be employed in the rubber industry. Because of its chem. indifference and highly voluminous structure cuprene presents a material with possibilities for a large no. of uses. Its application in the manuf. of acid resistant and elec. insulating materials, plastic masses, electrodes and explosives, as a cork substitute in the bookbind. industry and absorbing agent for illuminating materials and fuels or volatile matters is discussed.

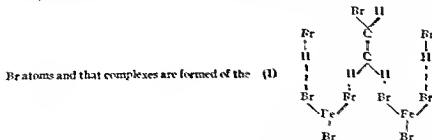
D. THURSEN

Nuclear synthesis in the olefin series. II. 1,4-Diolefins. BERNARD H. SNOWMAKER AND CLAUDE F. BOARD. *J. Am. Chem. Soc.* 53, 1505-12 (1931), cf. C. A. 24, 4756. By the use of CH_3CHCl_2MgBr the β Br ether synthesis of olefins has been extended to the prepn. of 1,4 diolefins. α Chloropropyl Et ether, b_p 34-6°, d_4^{20} 0.954, n_D^{20} 1.4120 (74% yield). α Chlorobutyl deriv., b_p 43-51°, d_4^{20} 0.938, n_D^{20} 1.4168 (81% yield). Bromination gives the $\alpha\beta$ -dibromoalkyl Et ethers, $RC(H)BrCH_2BrOPr$; R = Me, b_p 79-82°, d_4^{20} 1.467, n_D^{20} 1.5000, R = Et, b_p 99-101°, d_4^{20} 1.554, n_D^{20} 1.4968. With CH_3CHCl_2MgBr (details of its prepn. are given), there results α -allyl β -bromoalkyl Et ethers $RC(H)BrCH(C_2H_5)OL$, R = H, b_p 69-71°, b_m 84-6°, d_4^{20} 1.225, n_D^{20} 1.4592, R = Me, b_p 72-5°, d_4^{20} 1.162, n_D^{20} 1.4592, R = Et, b_p 88-92°, d_4^{20} 1.150, n_D^{20} 1.4600. By the action of Zn in $PrOH$ or $BuOH$ there results 1,4 pentadiene, hexadiene and heptadiene, the new compd. b_m 92-23°, d_4^{20} 0.7176, n_D^{20} 1.4202, the tetrabromide liquid, d_4^{20} 2.071, n_D^{20} 1.5734. The rule that unsatd. compds. with a double bond in the α position boil lower than the corresponding satd. hydrocarbons has been confirmed. It is further shown that in any given family the continuous chain diolefins form a definite series with gradually increasing b_p depending upon the relative position of the double bonds.

C. J. WEST

Addition of gaseous hydrogen bromide to acetylene and to gaseous vinyl bromide under the influence of catalysts and the addition of hydrogen bromide to vinyl bromide and allyl bromide in the liquid phase. J. I. WIRAUT. *Rec. trav. chim.* 50, 213-37 (1931), cf. C. A. 22, 1049, 24, 1076. An apparatus was designed wherein the rate of flow of the reacting gases could be maintained fairly const., in this way a series of expts. under well-defined conditions could be made. Otherwise conflicting results were obtained as to the proportion of isomeres formed, even in duplicate expts. The polymerization of vinyl bromide could be avoided by the addn. of 0.1-0.2% hydroquinone (cf. Mourou and Dufraisse C. A. 17, 2105). $MeCH_2Br$ was prepd. from $AcCl$ and PCl_5Br_2 (Perkin, *J. Chem. Soc.* 45, 522 (1884), Anschütz, *Ann.* 235, 300 (1886)), the contact substances consisting of metallic bromides or asbestos, were prepd. by impregnating the asbestos with the soln. of the bromide in concd. HBr , evapg. the latter on the water bath and drying in a current of dry HBr . Anhyd. HBr was prepd. by the action of Br on boiling tetrahydronaphthalene (Houben. *Methoden der organischen Chemie*, I 746) in an all glass app. The detn. of the solidification points of mixts. of $MeCH_2Br$ and $(CH_3)_2Br$ gave the following results: 0% $MeCH_2Br$, 99°, 34, 73°, 60, 54°, 10.6, 26°, 14.0, 6.6°, 20.8, -4.2°, 25.5, -3.2°, 32.8, -13.6°, 47.4, -26.0°, 57.8, -35.2°, 78.8, -61.0°, 84.6, -77.5°, 90.1, -76.5°, 94.4, -74.0°, 100, -63.5°. Samples contg. more than 60% $MeCH_2Br$ were analyzed by means of the n. 0% $MeCH_2Br$, n_D^{20} 1.5372, 10.0, 1.5122, 29.9, 1.5270, 49.9, 1.5214, 59.9, 1.5194, 69.9, 1.5167, 80.0, 1.5141, 84.9, 1.5129, 90.0, 1.5114, 100, 1.5080. In the expts. with vinyl bromide 50 g. was used in the course of 4 hrs., an excess of HBr being used; the results are given in the table below in the order: contact substance temp. (°C), di Br compds. (g), $(CH_3)_2Br$ in reaction product (%), empty tube 100, 0, . . .; glass wool 100, 59, 99.7, 100, 52, 100.0, 200, 6, 99.7, asbestos 100, 5, 68.1, silica gel 100, traces, 200, 18, 6, 100, 5, 3, asbestos from another lot 100, 42, 99.2, $HgBr_2$ on asbestos 100, 20, 4, 100, 18, 20, 290, 15.9, $HgBr_2$ cryst., no carrier 100, 0, . . .; $BiBr_3$ on asbestos 100, 16, 10, 100, 49, 49, 200, 19, 6, $BiBr_3$ powder, no carrier 100, 3, 3, 150, 15, 4, $SbBr_3$ on asbestos 100, 30, 41, 150, 22, 10, $AlBr_3$ on asbestos 100, 10, 48, 100, 34, 88.8, 200, 7, 0, 200, 6, 0, $AlBr_3$ powder, no carrier. black condensation product formed, $FeBr_3$ on asbestos 100, 32, 25%, 100, 66, 76, 200, 0, 33, $FeBr_3$ in small lumps, no carrier 100, 22, 9, 200, 6, 6. With acetylene the expts. were made in order to find out the conditions under which the dibromomethanes could be prepd. from C_2H_2 , twice the quantity of HBr required for the complete conversion of the C_2H_2 into dibromo-

ethane being used. Results with 6.41 acetylene and 26.01 HBr in the course of 4 hrs.: pure asbestos 100 traces, FeBr₃ on asbestos 100 traces, BiBr₃ on pumice 100 7, 65.6, 200, 13, 95.8, 200, 23, 94.5, 100, 9, 77.0, BiBr₃ on pumice 100, 2, 96.0, 200 1.5, 12 HgBr₂ on asbestos 100 23, 97, 150, 19, 94.4; 150, 29, 96.4 SbBr₃ on asbestos 100 10, 63, 100, 19, 87.1, 150 14, 56. It thus appears that the reaction between C₂H₂ and HBr is more difficult to realize than between vinyl bromide and HBr; nevertheless, the prepn of (CH₂Br)₂ from C₂H₂ and HBr is possible with reasonable yields. In the expts in the liquid phase 0.1 mol vinyl bromide was heated with 50% excess of HBr for several days at 100°. Another set of expts was performed with AcOH as a solvent and with the addn of metallic salts. The results are given in the following table in the ord r temp (°C), duration of heating, substance added, yield C₂H₃Br₂ (%), MeCHBr₂ in reaction product (%). Without solvent—100, 60 hrs. HBr d 1.60 to 100 100 66 hrs. HBr d 1.72 24 91, 100, 42 hrs. HBr d 1.82 40 69, 100 68 hrs. HBr d 1.89, 36, 65 H₂ 79% HBr in glacial AcOH—50, 14 days, none 32, 90, 50 14 days, none 32, 98, 50, 14 days, none, 37, 80, 50, 17 days, none 42, 76, 50 12 days, FeBr₃ 32, 97, 50 14 days, HgBr₂ 48, 100, 100, 44 hrs., none, 27, 70, 100 70 hrs. none 64, 71, 100 66 hrs. none 45, 75, 100, 66 hrs., FeBr₃ 53, 91, 100 71 hrs. HgBr₂ 42, 81, 100 68 hrs. AlBr₃ 50, 92, 100, 74 hrs., glass wool 48, 92, 100 47 hrs. asbestos 32, 97. In these expts MeCHBr₂ always is the chief product of the reaction while the metallic salts have little or no influence. With allyl bromide 15 g was heated with 42 g glacial AcOH and 24 g anhyd HBr in sealed tubes at 50° or 100° or was kept at room temp. either in the dark, in diffuse daylight or irradiated by an Uviol lamp. The highest yield obtained amounted to 48% and 1,2 dibromopentane always was the chief product of the reaction. It follows from these expts that the presence and nature of the solvent is the most important factor detg. the course of this reaction. If a liquid mixt of allyl bromide and anhyd HBr (without any solvent) is exposed to sunlight or to Uviol light (CH₂Br)₂ is the chief product of the reaction while in the dark a mixt of both isomers is formed. In the gas reaction between vinyl bromide and HBr glass wool and asbestos favor the formation of (CH₂Br)₂, while silica gel and ferric or Bi bromide, if used without a support, direct the reaction to the formation of MeCHBr₂. The reaction leading to MeCHBr₂ is the slower one of the 2 reactions, for in all reactions in which this isomer is obtained as the chief reaction product the yields are considerably lower than when (CH₂Br)₂ is the chief product. Obviously in a contact catalyst consisting of asbestos impregnated with a metallic bromide 2 catalytic activities are to be distinguished. With the freshly prepd contact substance the influence of the metallic bromide predominates and MeCHBr₂ is formed as the chief product. If, however, the metallic bromide has lost its catalytic properties, the influence of the carrier, the asbestos, predominates and the 1,2 isomer is the main product. The formation of the (CH₂Br)₂ is explained by the repelling action of the Br atom in vinyl bromide, while the influence of the catalytic properties of the metallic bromides is explained by the phenomenon of directed adsorption (Kruyt and van Duijn, C. A. 15, 3783, 22, 3321; 23, 5395). In the case of glass wool or asbestos there seems to be no reason why the mole. of vinyl bromide or HBr should be fixed on the adsorbent in a special manner, while for the contact substances, formed with metallic bromides, the assumption is made that the outer layer of the contact substance consists mainly of



type in which steric conditions are favorable for the formation of MeCHBr₂. The results obtained with vinyl and allyl bromide in the liquid phase are considered not to be in accordance with the theory of alternating polarity. C. F. VAN DUYN

Oxidation of methanol with air over iron, molybdenum and iron-molybdenum oxides. HOMER ADKINS AND WESLEY R. PETERSON. J. Am. Chem. Soc. 53, 1512-20 (1931) —

of 2-ethoxy-1 butene, b. 55-7°, n_D^{20} 1.4011. PhMgBr gives 60% of α -phenyl- β -bromoethyl Et ether, b. 129-33° n_D^{25} 1.5317. KOH gives 62% of $\text{CH}_3\text{C}(\text{OEt})\text{Ph}$, b. 109-12°, n_D^{25} 1.5287. $\text{MeCHBrCH}(\text{OMe})\text{Br}$ and PhMgBr give 71% of $\text{MeCHBrCH}(\text{OMe})\text{Ph}$, b. 122-4°, n_D^{25} 1.5316. KOH gives 42% of α -methoxy β -methylstyrene, b. 96-8°, n_D^{25} 1.5271. $\text{tC}_4\text{H}_9\text{BrCH}(\text{OMe})\text{Br}$ and tMgBr give 56% of 3-bromo-4-methoxyhexane, b. 74-8°, n_D^{25} 1.4405. KOH gives 79% of 3-methoxy-4-hexene, b. 114-5°, n_D^{25} 1.4130. These unsatd ethers in the presence of a trace of acid or even semicarbazide- HCl , hydrolyze to the corresponding ketones. C. J. WEST

Methyl selenomercaptan. A. BARONI *Atti accad. Lincei* 12, 234-7 (1930).—MeSeH was prepd by treating alc NaSeH with MeI. The NaSeH was made by satg a 10% soln of $\text{I}(\text{ONa})$ with H_2S until the Na-S first pptd redissolved. The reaction was carried out in a flask equipped with a condenser, the receiving flask was cooled with liquid CO_2 and Me_2CO the product being distd off by heating on a water bath. MeSeH is an almost colorless liquid having a very disagreeable odor, b.p. 12°. The following metal salts were also prepd: $(\text{CH}_3\text{Se})_2\text{Hg}$ is formed by addn of $\text{Hg}(\text{CN})_2$ while cooling strongly as yellow crystals. The Pb salt is obtained by adding Pb(OAc)₂, the Ag salt by means of AgNO₃ and NaOAc, the Cu salt with Cu(OAc)₂, and the Bi salt with BiCl₃ and NaOAc. These are all obtained as yellow cryst. powders, whose analysis agreed very closely with that calcd. A. W. COVIER

Effects of activated and non-activated magnesium-copper alloy on the yields of some organomagnesium halides. HENRY GILMAN AND E. A. ZOELLNER. *J. Am. Chem. Soc.* 53, 1581-3 (1931).—When equiv quantities of Cu-Mg alloys are used alone, either in the activated or non activated condition, the yields of some Grignard reagents are lowered. However, small but effective quantities of the activated alloy appear to have no significant effect on the yield. The other more commonly used catalyst, I, apparently has the same effect as the alloy on the yields of some Grignard reagents. C. J. WEST

Preparation of organomagnesium halides in the presence of magnesium iodide. Studies on the capture of free radicals. HENRY GILMAN AND E. A. ZOELLNER. *J. Am. Chem. Soc.* 53, 1583-6 (1931).—In studies concerned with the capture of free radicals formed during the prepn of organomagnesium halides it has been shown that the quantity of I used to catalyze the formation of *tert* BuMgI has a significant effect on the yield. There is no appreciable lowering of yield with a very small quantity or a full atom equiv of I. Another catalyst, the Mg-Cu alloy which is used with more refractory RX compds., also appears to have a corresponding effect on the yields of other RMgX compds. The av. % yields of Me_2CMgCl with varying amts of I (at. equiv given) are 0, 83.3, 0.0125, 83.6, 0.023, 77.1, 0.05, 72.2, 0.1, 63.5, 0.2, 64.9, 0.3, 64.8, 0.4, 71.8, 0.6, 76.2, 1.0, 79.7. C. J. WEST

Some mechanisms for the formation of organomagnesium and other organometallic compounds. HENRY GILMAN AND R. E. BROWN. *Rec. trav. chim.* 50, 184-7 (1931).—The formation of Grignard reagents may be explained by the reactions: $\text{RX} + \text{Mg} \rightarrow \text{R}^- + -\text{MgX}^+$ (I) $\text{R}^- + -\text{MgX}^+ \rightarrow \text{RMgX}$ (II). This mechanism, however, does not include all organomagnesium compds., e. g., the prepn of MgEt_2 from H_2Et_2 and Mg; moreover, the Mg subhalides are hypothetical and the possibility exists that the reactions of the MgX compds. may be essentially those involving a catalytic effect of MgX_2 on Mg (cf Gilman and Brown, *C. A.* 24, 1837). Therefore, the following reactions are proposed for the formation of organomagnesium compds.: $2 \text{RX} + \text{Mg} \rightarrow 2 \text{R}^- + \text{MgX}_2$ (III), $2 \text{R}^- + \text{Mg} \rightarrow \text{R}_2\text{Mg}$ (IV), $\text{R}_2\text{Mg} + \text{MgX}_2 \rightarrow 2 \text{RMgX}$ (V). There is definite exptl evidence in support of reaction III, e. g., the transitory existence of benzoyldryl radicals (Gilman and Zoellner, *C. A.* 24, 574u) while there is also very compelling evidence for the correctness of reaction V (Gilman and Fothergill, *C. A.* 23, 546S; Gilman and Brown, *C. A.* 24, 1833). All that is needed is support for reaction IV and this will probably come forth from the work of Paneth on free Me and Et radicals (*C. A.* 23, 5159, 24, 3750), in which it is shown that some organometallic compds. decompose by pyrolysis *in vacuo* into the formation of the metal and free radicals: $\text{Et}_2\text{Pb} \rightarrow 4 \text{Et}^- + \text{Pb}$, which can then unite with some other metals to give other organometallic compds. Although no expts. were carried out with Mg, the close relationship between Zn and Mg suggests the possibility of such reactions with Mg. Expts. carried out with Et radicals, obtained by Paneth's method from PbEt_4 and a Be mirror, did not result in the formation of BeEt_2 and thus it could not be decided whether the removal of the Be mirror is due to atomic H such as had been suggested by Schultze and Muller (*C. A.* 24, 2036). C. F. VAN DUSEN

Effect of copper upon the yields of Grignard reagents. GEO. JOHNSON AND HOWARD ANKIN. *J. Am. Chem. Soc.* 53, 1530-3 (1931).—The following yields were obtained with Mg and a Mg (88%) Cu (12%) alloy. BuBr 92.5, 74.7, sec BuBr 73.8, 49.6, BuI 82.4, 64.2, EtI 82.4, 76.1, PhCH₂Cl 94.1, 80.2, cyclohexyl bromide 82.5, 64.7, allyl bromide 72.1, 6.0, PhBr 90.9, 89.1. These values show that the yields of the Grignard reagent were very considerably lowered, except from PhBr, by the presence of Cu, which promoted the competitive Wurtz reaction. C. J. WEST

Organic lead compounds. I. Action of acids on lead aryls. PAUL R. ALSTINE. *J. Am. Chem. Soc.* 53, 1548-52 (1931).—The strong tendency of organo-Pb compounds to retain at least 2 Pb-C linkages has been further demonstrated by the action of HNO₃ and HBr on *tert*-Bu-Me₂Pb, *tert*-Bu₂Pb, and *tert*-Bu₃Pb. Heating (a) Me₂C(CH₃)₂PbH in C₆H₅Me for 24 hrs. gives *tert*-butyllead bromide m 201.2 (a) Me₂C(CH₃)₂PbBr (18 g) and PhMeBr give 16 g phenyllead bromide m 181.2 (a) Me₂C(CH₃)₂PbH (8 g) and 40% HBr in CHCl₃, shaken for 2 hrs. give 1 g *tert*-butyllead bromide m 129-30°. The action of a larger quantity of HBr or the action of HBr in CHCl₃ on (a) Me₂C(CH₃)₂PbH gives *di-tert*-butyllead bromide m 140.1 (a) Me₂C(CH₃)₂PbPh and 40% HBr, refluxed for 45 min., give 1 g of phenyllead bromide m 117-8°. A larger quantity of HBr, refluxed 4 min. gives 44% of phenyllead bromide m 116-7°. (a) Me₂C(CH₃)₂PbPh and HNO₃ in petrol ether give 69-93% of phenyllead bromide dihydrate; the same product was obtained even with boiling HNO₃ (d 1.42), together with some *di-tert*-butyllead dihydrate. A new volumetric method for the detn. of Pb in aryl Pb compounds described. C. J. WEST

Reactions of some organic lead compounds. CHARLES D. HURN AND PAUL R. ALSTINE. *J. Am. Chem. Soc.* 53, 1547-8 (1931).—PhEt₂Pb and EtCl give C₆H₅Et and Et₂PbCl further reaction gives the hexavalent PbCl₆. Both PhEt₂Pb and Et₂Pb react readily with HNO₃ to produce Et₂Pb(NO₃)₂, which reacts metaethetically in MeOH with HCl to give the dichloride. In 0.1 M aq. soln. the dihydrate is apparently 70% ionized. Pyrolysis of Pb(OEt)₂ gives rise to EtOH and Pb dibenzate, m 119-20° (crystd. from MeOH) or 145-7° (block, after crystn. from C₆H₅Cl). Similarly, Pb(O₂C₂H₅)₂ gives AcOH and Pb(O₂C₂H₅). C. J. WEST

Synthesis of formaldehyde by catalytic oxidation of methanol. H. KADOWAKI AND M. YAMADA. *Repts. Osaka Imp. Ind. Research Lab. Japan* 11, No. 19 (1931), d. C. A. 23, 4020.—A Ag catalyst was prepd by evap. to dryness a mixt. of AgNO₃ (2.5-3.0 g) and pumice (1 g) in water and later heating to 400-600° in an elec. furnace. The expt. was carried out under the same conditions as those reported previously. The effects of other elements are also discussed: Fe, Co, Cr, Mo, W, U, Mn, Zr, Th, Cu, Zn and Al show poisonous effects while Se and V slightly accelerate the reaction. Excess CO₂ is harmful, but its effect is negative if the % of CO₂ in the air is below 8%. I. I. NAKAMURA

The oxidation of gaseous acetaldehyde with oxygen as an example of the combustion of hydrocarbons. MAX BODENSTEIN. *Sitzb. preuss. Akad. Wiss.* 1931, 73-88.—The reaction between gaseous AcH and O₂ at 0.1 atm. pressure and 70° gives Ac₂O₂H as the principal product. The rate is given by $d(\text{Ac}_2\text{O}_2\text{H})/dt = k(\text{AcH})^m(\text{O}_2) + n$, where $m = 1.50$ if the pressure is measured in mm. bromonaphthalene. It is not a chain reaction but the mechanism is as follows: A mol. of AcH becomes activated, reacts with O₂ and forms an energy rich mol. of Ac₂O₂H. This transfers its energy to the AcH mol. primarily but secondary reactions are with O₂, negative catalysts or with the walls. The results are discussed in connection with the oxidation of C₂H₂ and other gaseous hydrocarbons. It is probable that the scheme could account for the oxidation of all hydrocarbons. G. M. MURPHY

The course of the reaction between citronellal oxime and acetic anhydride. J. ZIMMERMANN. *Rec. trav. chim.* 50, 283-6 (1931).—On boiling with Ac₂O, citronellal oxime passes into citronellonitrile (Semmler. *Ber.* 26, 2235 (1893)) and on this reaction the method of Dupont and Labaune (*Ber. Route Bertrand fils* 1912, 3) for the quant. detn. of citronellal in essential oils is based. After having boiled the oxime with Ac₂O, the reaction product is treated with 0.5 N alic. KOH, the assumption being made that this has no influence on the nitrile. This assumption is, however, invalid, 7-12% of a non volatile residue being obtained by Z. in this way. The reaction between citronellal oxime and Ac₂O is a very violent one if NaOAc is present, within a short time the temp. rises to 145°, while without NaOAc this violent reaction only occurs after heating on the water bath. In this reaction a partial resinification takes place and the formation of citronellamide besides a small amt. of a N compd. with a cone. like odor was obtained, the same reaction occurs with concd. H₂SO₄ and, therefore, a N detn. cannot be carried out by the Kjeldahl method. The citronellonitrile formation

is therefore unsuitable for the quant detn of citronellal. The citronellal used in these expts was purified through the bisulfite compd; it b_p 68–71°, d₄²⁷ 0.8435, n_D²⁷ 1.4435, α_D²⁵ 10.10', citronellal oxime, b_p 124–5°, d₄²⁷ 0.8971, n_D²⁷ 1.4711.

C. F. VAN DUIN

The action of Japanese acid clay on terpene compounds. II. Cyclization of citronellal. TSUTOMU KUWATA. *J Soc Chem Ind, Japan* 34, Suppl binding, 70–2(1931); cf *C A* 24, 1857.—This clay with citronellal in MeOH gives 32% isopulegol and compds contg 23.44% MeO groups. No menthogycol was found.

V. F. HARRINGTON

Syntheses from natural gas hydrocarbons. I. Caproic acid from pentane. H. B. HASS AND J. R. MARSHALL. *Ind Eng Chem* 23, 352–3(1931).—Treatment of AmCl with NaCN gave 70% yields of AmCN, with recovery of 25–28% of unchanged chloride. Other primary chlorides also reacted satisfactorily, but secondary chlorides or bromides gave poor yields (approx 30%), and tertiary chlorides and bromides, no nitrile. NaI, acting as an intermediate compd catalyst, leads to 90% yields, although its use is only recommended for cyaniding the less reactive chlorides and bromides. Hydrolysis of the AmCN with 67% H₂SO₄ is complete in 30 min. Sepn of the nitrile from the halide is possible because AmCl and BuCl form binary min-boiling azeotropic mixts with EtOH and ternary azeotropic mixts with EtOH and H₂O. These processes are suitable for the manuf of caproic acid with EtOH and ternary azeotropic mixts. The method is a general one for the conversion of any primary alkyl chloride to the fatty acid contg 1 more C atom.

ALBERT THOMAS FELLOWS

The bromides of the eleostearic acids. I. The hexabromides. J. VAN LOON. *Rec trav chim* 50, 32–6(1931).—Recent researches have shown the eleostearic acids have the formula Me(CH₂)₅CH=CHCH=CHCH=CH(CH₂)₅CO₂H (Libner and Rossmann, *C A* 24, 4839, Böseken, *C A* 23, 5160, van Loon, *Diss Delft*, 1929, 62, *C A* 24, 4178). The formation of hexahalogen compds, however, does not take place easily and Kaufmann was the first to show the presence of 3 double bonds (*C A* 20, 2989) tetrametrically although he did not isolate any hexa Br compd. In order to prep these compds β-eleostearic acid was brominated in CCl₄ with the theoretical amt of Br under the influence of Uviol light after it had been shown that this light does not provoke any particular intramol change. After evapn of the solvent and recrystn of the residue from petroleum ether to which a small amt of ether had been added, 0.5 g of the hexabromide, m 157°, was obtained from 7 g of the acid. On debromination with Zn and alc. HCl the hexabromide gave β-eleostearic acid in quant yield, the same acid was again obtained on debrominating the oily mixt. of hexabromides which results after sepn of the cryst compd. The same cryst hexabromide was also obtained from α-eleostearic acid, which is easily converted into the β form by traces of balogen, therefore most probably the hexabromides are formed in both cases from the same mixt of tetrabromides which are formed as the primary reaction products, tetrabromination taking place instantaneously and further bromination only very slowly. The solid tetrabromide, m 114°, on bromination gives the same hexabromide, m 157°, the yield being only 10 g from 100 g of the tetrabromide, in this case again the oily hexabromides on debromination gave β-eleostearic acid. C. F. VAN DUIN

Preparation of dichloroacetic acid. HOWARD W. DOUGHTY AND GERHARD J. DERGE. *J Am Chem Soc* 53, 1591–6(1931).—A simple and convenient method is described for the prepn of Cl₂CHCO₂H by the reaction of Cu with Cl₂CCO₂H in H₂O. The hydrolysis of Cl₂CCO₂Cl has been studied at 60° and 100° and the temp coeff. found to be 2.576. The yield of the acid, m 5–6°, d₄²² 1.5594, was about 80%.

C. J. WEST

Esterification in the presence of anhydrous salts. FRANCIS M. WHITACKER AND H. T. BRISCOE. *Proc Indiana Acad Sci* 38, 187–94(1929).—The anhyd sulfates of Cu, Cd, Al, Mn, Fe, Ca and K alum were tested as catalysts in the esterification reaction between EtOH and HOAc. Higher alcs and higher acids and branched-chain compds were used with CuSO₄. CuSO₄ gave the highest yield of EtOAc, 84.1% after 264 hrs refluxing. Plaster of Paris was a retardant and anhyd Na₂SO₄ had no effect. The temp at which the salt is dehydrated is important. When the temp at which the salt is dehydrated is high enough to remove all the water of crystn the promoting effect is accelerated but this accelerating effect is reduced when the temp of dehydration is high. W. and B. conclude that the promoting effect of anhyd salts is partly due to removal of water and partly to other causes.

H. M. STARK

Esterification with thiolacetic acid. FRANCIS B. STEWART AND PAUL V. McKINNEY. *J Am Chem Soc* 53, 1482–90(1931).—Me·CHOH and AcSH yield only H₂S and the normal ester, Ph₃COH only H₂O and the thiol ester, MeOH and EtOH

give 78% H_2S and 22% H_2O . $MeOH$ appears to esterify somewhat more rapidly than $EtOH$ but both are much more reactive than Me_2CHOH , the data for Ph_3COH are at 100° and since the higher temp. was necessary to obtain accurate results it is unquestionably less reactive than Me_2CHOH . The quantities of H_2S evolved by the 3 reactions show that there are 3 distinct processes. In Ph_3COH only the H_2O group enters into the reaction, in Me_2CHOH only the H takes part in the reaction, while in the cases of $MeOH$ and $EtOH$ either H or OH can split off. The question of the mechanism of esterification in general is considered and an explanation of certain anomalies is attempted. C. J. Wray

Preparation of diacetamide hydrochloride and its use as an acetylating agent. HENRY STEPHEN. *J. Chem. Soc.* 1931, 4723. $AcNH_2$ and $SOCl_2$ in 1:1 ratio give $(AcNH_2)_2HCl$ and $MeCN$. When $(AcNH_2)_2HCl$ is fused with an equimol. proportion of the HCl of a primary aromatic amine the Ac deriv. of the latter is formed together with a trace of an amide. The reagents have been applied with success to the prepn. of $PhNHAc$, o - m - and p - MeC_6H_4NHAc and α - and β - C_6H_5NHAc . C. J. Wray

The preparation of glycol chlorohydrin. E. D. G. FRANK. *Rec. trav. chim.* 50, 261 (1931). $HOCH_2CH_2Cl$ (I) is always prepd. by the addn. of HCl to C_2H_4 , the HCl being used in the form of $Ca(OCl)_2$ (cf. Norms, C. A. 13, 2749), $NaOCl$ (1 sees and Ward U. S. Pat. 1,534,605, C. A. 26, 3179) or Cl_2 and water (Gomberg, C. A. 13, 2669). Irvine and Haworth (U. S. Pat. 1,496,675, C. A. 18, 2341) have proposed to use a 5% $CuCl_2$ soln. as a catalyst in this process but their method of prepn. is far too complicated. It is now pointed out that the reaction proceeds better on using 1% $CuCl_2$, while with 5% $CuCl_2$ less favorable results are obtained, the best results are, however, obtained with 0.1% $CuCl_2$, a rapid absorption of the C_2H_4 taking place with hardly any formation of $(CH_2Cl)_2$. Without the addn. of Cu salts a high yield of I may only be obtained when an excess of C_2H_4 is present during the whole course of the reaction; under these conditions, however, losses of C_2H_4 cannot be avoided. The formation of I is not only catalyzed by Cu salts, but Ni , Co and Fe salts exert the same action and it is noteworthy that all these salts accelerate the decompos. of HCl into H_2 and O as well. These results show that the formation of I from C_2H_4 , water and Cl_2 is not to be considered as an addn. of HCl to C_2H_4 , but as an oxidation of ethylene oxide, followed by a reaction of the latter compd. with HCl to form I. This conception of the reaction is supported by the formation of both isobutylene oxide and Me_3COH on passing isobutylene into I HCl (Pogorelec, *Chem. Zentr.* 1905, 1, 797). The amt. of acid formed in the reaction always surpasses the yield of I, which points to the formation of higher chlorinated products. By distn. of a larger amt. of $(CH_2Cl)_2$, obtained in several expts., a small amt. of tetrachloro- and pentachloroethane could be isolated. C. F. VAN DUSEN

Reaction between propylene oxide and diethylamine. K. A. KRASUSHIN AND G. T. PIVTOLAT. *Ukrainski Khim. Zhur.* 5, Ser. Pt. 135-6 (1930). The reaction between 1 mol. of $MeCH_2CH_2O$ and 2 mols. of Et_2NH in water at concns. of 33 or

70% results in the formation of 65% of 3-diethylamino-2-propanol (I), d_4^{20} 1.575-9°, d_4^{20} 0.8677, d_4^{20} 0.8511. The HCl salt is prepd. by passing dry HCl into I in Et_2O and recrystg. from Me_2CO . The picrate is obtained by mixing I and picric acid in mol. proportions in alc., m 89°. CHAS. BLANC

The reaction between ethylamine and diethylamine and isopropylethylene oxide. F. P. KRIKOVOS. *Ukrainski Khim. Zhur.* 5, Ser. Pt. 141-6 (1930). The reaction between isopropylethylene oxide, $Me_2CHCH_2CH_2O$ (II), and Et_2NH gives α -ethyl

amino- β -hydroxy- γ -methylbutane (III) when Et_2NH is used in excess (2 mols. of amine to 1 mol. of I), while excess of I (1.5 mols. of I to 1 mol. of amine) produces β -hydroxy- γ -methylbutyl-diethylamine (III). Similar reaction between I and Et_2NH in all proportions gives only α -diethylamino- β -hydroxy- γ -methylbutane (IV). II was prepd. when a mixt. of 10 g. of I and 37% aq. Et_2NH was heated 10 hrs. in a sealed tube in a water bath, the reaction product treated with K_2CO_3 , s.p.d., dried with KOH and redistd., yield 8 g. d_4^{20} 1.50-2°, d_4^{20} 0.9076, d_4^{20} 0.8929. The HCl salt, m 145°, is obtained by passing dry HCl into II in Et_2O . The picrate, m 119-56°, results by mixing the alc. solns. With an excess of I under similar conditions are obtained 5 g. of II and 3.5 g. of III, d_4^{20} 1.244-7°, the HCl salt is a heavy liquid. IV is prepd. when a mixt. of 15 g. of I and 15 g. of Et_2NH in 50 g. of water is treated as described above, yield 14 g., d_4^{20} 1.51-4°, d_4^{20} 0.8671, d_4^{20} 0.8512. The HCl salt is formed by dropping HCl into IV in water to acid reaction, evapg. and allowing to stand in a vacuum desiccator. The

picrate, m 80–1°, was obtained by mixing equimol amts of alc solns. of the acid and IV, adding drop by drop water to permanent turbidity, allowing to stand and filtering.

CHAS BLANC

Tin tetrahalides and dioxane. HEINRICH RHEINBOLDT AND RICHARD BOY. *J. prakt. Chem.* 129, 268–72 (1931). — $O(CH_2CH_2)_2O$ and $SnCl_4$, with cooling, give the compd $SnCl_4 \cdot 2C_4H_{10}O_2$, sol in cold H_2O and insol in CS_2 , the same compd is prepd by soln of $SnCl_4$ in Et_2O . $SnCl_4 \cdot 2BzH$, $SnCl_4 \cdot 2BzMe$ or $SnCl_4 \cdot 2PhCH_2CH_2CH_3$ in dioxane. The compds $SnCl_4 \cdot 2C_4H_{10}O_2$ and $SnI_4 \cdot 2C_4H_{10}O_2$ were also prepd. The dark violet iodide turns yellow in 2–3 hrs.

C J WEST

Dioxane and halogens. HEINRICH RHEINBOLDT AND RICHARD BOY. *J. prakt. Chem.* 129, 273–7 (1931). — Wurtz (*Ann. chim. phys.* 69, 321 (1863)) described the compd $O(CH_2CH_2)_2O$, OBr_2 . Kehrman and Falke (*C. A.* 19, 816) prepd the same compd, but ascribed to it the formula $O(CH_2CH_2)_2O \cdot 2HBr$. The following work confirms Wurtz's structure. The compd was prepd from 10 cc dioxane and 5 g Br with or without cooling. The Br detd titrimetrically and gravimetrically was the same, thus confirming the Wurtz structure, the compd, orange, m 65–6°, is stable for only a short time, is not attacked by cold H_2O but is easily decompd by alkalis. I gives the compd $C_4H_8O_2$, I, red violet, m 81.5°, and is more stable than the Br compd, I being split off only after several hrs. ICl gives the compd $C_4H_8O_2$, ICl, red-brown, m 50–8°, and is unstable. The compd $C_4H_8O_2$, IBr, red brown, m 65°, is stable for only a short time.

C J WEST

The methylation of putrescine. WERNER KEIL. *Z. physiol. Chem.* 196, 81–6 (1931). — Of the 9 possible methylation products of putrescine 7 are already known. Four of these have again been prepd, also the 2 missing members, thus completing the series. N,N,N',N' -Tetramethylputrescine, b 167–70°, was obtained by heating $(CH_3)_2(NH)_2$, 2HCl with paraformaldehyde in an oil bath at 100–210° for 0.5 hr., then adding H_2O , filtering off the excess paraformaldehyde, adding NaOH to the filtrate and extg with $CHCl_3$, chloroaurate, m 210°, chloroplatinate, m 230–7°. N,N' -Dimethylputrescine was prepd by heating $(CH_3)_2(NH)_2$, 2HCl with 40% CH_3OH and $CaCl_2$ until the temp reached 170°, adding H_2O and NaOH, filtering off the $Ca(OH)_2$, and extg the filtrate with $CHCl_3$, chloroaurate, m 172°, chloroplatinate, m 237°. Under the same treatment $Me_3N(CH_2)_3NH_2$, 2HCl yielded N,N,N' -trimethylputrescine, chloroaurate, m 151°. Methylation of $PhO(CH_2)_3NH_2$ with Me_2SO_4 gave δ -phenoxymethyl- N,N,N' -trimethyltetramethyleammonium chloride (chloroaurate, m 112°), which was then refluxed with HI. The resulting δ -iodo- N,N,N' -trimethyltetramethyleammonium iodide (I), after removal of $PhOH$ by Et_2O , was heated with alc $MeNH_2$, 1 hr at 100°. Alc and $MeNH_2$ were removed by evapn, HI by $TiOAc$ and excess Ti by HCl. The product was δ - N' -methylamino- N,N,N' -trimethyltetramethyleammonium chloride (II), corresponding to the base $MeNH(CH_2)_3NMe_3OH$, chloroaurate, m 205–6°. A similar condensation of I with Me_2NH yielded δ - N,N' -dimethylamino- N,N,N' -trimethyltetramethyleammonium chloride (III), corresponding to the base $Me_2N(CH_2)_3NMe_3OH$, chloroaurate, m 244–5° (decompn), picrate, m 260°. N,N' -Dimethylputrescine was obtained by heating $Cl(CH_2)_3N(CO)_2C_6H_5$ with Me_2NH in $EtOH$, evapn to dryness, adding KOH and distg. to dryness several times, the chloroaurate was prepd as usual. II and III are the new methylated putrescines derivs. A. W. DOX.

The structure of the "mercaptomethanesulfonic acid" (thiosulfuricmethionic acid), H J BACKER. *Rec. trav. chim.* 50, 268–78 (1931), cf *C. A.* 25, 76. — The acid formed by the interaction of Cl_3CSCl and K_2SO_3 (Rathke and Albrecht, *Ber.* 3, 358 (1870), *Ann.* 161, 129 (1872)) has, up to the present, always been considered to be $(HO_2S)_2CSH$, on oxidation with Br this acid gives $CH(SO_3H)_2$, exactly 6 atoms of Br being necessary. $(KO_2S)_2CSH + 6Br + 4H_2O \rightarrow (SO_3H)_2CH + 6HBr + H_2SO_4$ (cf Backer, *C. A.* 25, 76). The acid, however, does not form mercaptides, reacts neutral, does not give ppt with Pb, Cu and Hg salts, is odorless and dissolves HgO with the formation of a white ppt, and the absence of most of the characteristic mercaptan reactions rendered the constitution as a mercapto compd doubtful. On adding alkali to the soln of the acid, a yellow color is developed, which disappears on acidification, subsequent addn of $FeCl_3$ then produces a dark blue color. This color reaction with $FeCl_3$ is obtained only when alkali has first been added and thus is not a characteristic reaction for the "mercapto compd" itself, but for the degradation product obtained by the action of alkali. The same series of reactions is given also by thiosulfuric-carboxylic acids, their salts and esters which also dissolve HgO with the formation of a white ppt. On heating salts of the acid, SO_3 is evolved with the formation of sulfate, the same reaction taking place on heating the soln with HCl, with KCN a ppt of $BaSO_4$ is formed if Ba salts are present. All these facts and analogies point

Recent work by Hill and Hibbert (*C A* 18, 1937) has shown that glyceryl acetal made from paraldehyde consists of 80% I and 20% II, and that from the C_3H_8 method ($HgSO_4$ and H_2SO_4 catalyst) is 64% I and 36% II. Ethylideneglycerol was made by N. and F. from C_3H_8 , using $HgSO_4$ and anhyd fluo acids as catalysts, and the low temp of 37° . The products consisted of 78% of the I and 22% of II. The higher yield of the I in this case is attributed to the use of the fluo acid catalyst and to keeping the liquid cool during the reaction.

H. M. STARK

Polymerization of methyl esters of highly unsaturated acids. VI. Hydrogenation of the polymerized product. VII. Hydrogenation of linseed oil and of the methyl esters of its liquid acids. KICHIRO KINO *Sci Papers Inst Phys-Chem. Research* (Tokyo) 15, 127-9, 130-6 (1931) (in German), cf *C A* 25, 914—The contentious question is reexamd, whether polymerized acids suffer depolymerization when hydrogenated. The intrapolymerized complex $C_{41}H_{72}CO_2Me$ (3 double bonds), was hydrogenated in presence of Pd in AcOH at 90° , and without solvent at $180-220^\circ$. The I no was strongly reduced, the mol wt remained const. In AcOH, more II, was absorbed than could be calcd from the I nos. It is assumed, therefore, that the 4-C ring was reopened. Linseed oil, I no 183, sapon no 194, mol wt, 886, was polymerized by heating it under II, for 5 hrs at $290-300^\circ$. The treated oil was extd 3 times with 3.5 vols. of Me_2CO . The first Me_2CO ext (I), and the final insol residue (II) showed I no 118, 104, sapon no 191, 195, mol wt 861, 2311. The liberated acids from I and II were distd in *vacuo* to sep the liquid monomol acids from higher polymers. I furnished mainly intrapolymerized acids (III), II gave largely nondistillable high polymers (IV). The acids from III and IV, purified through the Pb salts, were hydrogenated in AcOH, in presence of Pd, at $50-70^\circ$. The absorbed H_2 corresponded nearly to the figures calcd from the I nos. Characteristics of the acids from III, before and after hydrogenation, were I no 146, 65 (calcd from II, absorption, 9), mol wt 283, 278. The high I no of the hydrogenated product must be ascribed to reopening of the intramol 4 C ring. The hydrogenated acids from IV showed unchanged mol wts (564, 568). Hydrogenation, therefore, did not cause depolymerization of intermol combined acids. A similar examn of the Me esters of the liquid acids from linseed oil led to the same conclusions.

K. H. ENGL

Continued studies on the splitting off of halogen from haloacylamino acids and polypeptides by dilute alkali. EMIL ANGERHALDEN and EDGAR HAASE *Fermentforschung* 12, 313-28 (1931)—The rate at which halogen is split off by dil alkali varies considerably with different haloacylamino acids according to the acyl group which carries the halogen and the configuration of the amino acid to which the acyl is attached. A series of $BrCH_2CO$ derivs has now been prepd and studied with respect to rate of halogen cleavage. The derivs were obtained by treating the amino acid or peptide soln in *N* NaOH with $BrCH_2COBr$, and comprise the following bromoacetyl glycine (I), m. $115-6^\circ$; bromoacetyl α -alanine (II), m. $96-7^\circ$; bromoacetyl β -alanine (III), m. $119-20^\circ$; bromoacetyl α -aminobutyric acid (IV), m. $159-60^\circ$; bromoacetyl β -aminobutyric acid (V), m. $99-100^\circ$; bromoacetyl α -valine (VI), m. $142-4^\circ$; bromoacetyl β -aspartic acid (VII), m. $134-5^\circ$; bromoacetyl γ -glutamic acid (VIII), sirup, bromoacetyl β -proline (IX), m. $118-20^\circ$; bromoacetyl β -phenylalanine (X), m. $116-9^\circ$; bromoacetyl β -tryptophan (XI), m. $158-9^\circ$; di(bromoacetyl) β -cystine (XII), m. $126-7^\circ$; bromoacetyl glycidyl β -valine (XIII), m. $170-1^\circ$; bromoacetyl β -valylglycine (XIV), m. $155-6^\circ$; bromoacetyl glycidyl β -leucine (XV), m. $145-6^\circ$; bromoacetyl β -leucylglycine (XVI), m. $147-8^\circ$; bromoacetyl β -alanyl β -leucine (XVII), m. $154-5^\circ$; bromoacetyl glycidyl β -alanyl β -leucine (XVIII), m. $167-8^\circ$; \rightarrow diglycidyl β -alanyl β -leucine, m. $110-2^\circ$ (decompn). \rightarrow bromoacetyl glycidyl β -alanyl β -leucine (XIX), m. $168-70^\circ$ (decompn). These haloacyl derivs were subjected to the action of NaOH at pH 8.4 and 37° . The liberated Br ion was detd. by the Volhard titration at intervals of several hrs. By far the most rapid liberation of Br was from IX. The least reactive was VII, then in increasing order, XI, X, VIII, bromoacetyl β -tyrosine, VI, III, bromoacetyl β -leucine, I, V, II and IV. Observations were not conclusive with XII because of its insol. Among the dipeptide derivs XIII gives up Br more readily than its isomer XIV, and XVI more readily than its isomer XV. A comparison of the $BrCH_2CO$ derivs of di-, tri-, and tetrapeptides, XVII, XVIII and XIX, shows that the rate of Br cleavage increases with increasing length of the peptide chain. The liberation of Br from I was not greater in the presence of 0.5 *N* H_2SO_4 than in H_2O alone.

A. W. DOX

The behavior of dipeptides containing *m*- and *o*-tyrosine, *p*-nitrotyrosine and phenyl- β -alanine toward N alkali, erepsin and trypsin-kinase. EMIL ANGERHALDEN and WALTER SCHÄRRER. *Fermentforschung* 12, 295-312 (1931)—Of the dipeptide derivs contg. no naturally occurring component, chloroacetyl β -*o*-tyrosine is the only one thus far found

to the ester by enzymic hydrolysis. 4 no. of new derivatives have now been prepd containing glycine or leucine or the corresponding haloacetyl amino acids, linked to *m*- and *p*-tyrosine, *o*-nitrotyrosine and phenyl β -alanine. The tyrosines were prepd. by previously described methods. Glycine anhydride was condensed with *o*-HOCH₂CHO, NaOAc and Et₃Ac to give *o*-acetoxycarbonyl-2,5-diketopiperazine, the latter reduced and hydrolyzed (I) and (II), and the resulting *dl*-*o*-tyrosine purified by esterification and upon hydrolysis the ester *m*-HOCH₂CHO was condensed with hippuric acid, NaOAc and AcO to the acetylated lactamide of *m*-hydroxybenzoylaminotyrosine, the latter hydrolyzed and reduced to *m*-benzoyltyrosine and finally the Br removed by hydrolysis to yield *dl* *m*-tyrosine. Direct nitration of ordinary tyrosine gave 2-nitro-*p*-tyrosine. The coupling of these tyrosines with haloacetyl amino acids and amidation to the dipeptides was performed in the usual manner. 1 Tyrosine + ClCH₂COCl \rightarrow chloroacetyltyrosine \rightarrow glycyltyrosine decomps. 130°. 1 tyrosine ester + *dl*- α -Me-CHCH-CHBrCOBr \rightarrow *dl*- α -bromosuccinyltyrosine ester, m 76°, \rightarrow free acid \rightarrow *dl*-leucyltyrosine + 2 H₂O. *dl*-*o*-tyrosine Et ester + ClCH₂COCl \rightarrow *dl*- α -bromosuccinyl-*dl*-*o*-tyrosine Et ester, mters 50°. \rightarrow chloroacetyl-*dl*-*o*-tyrosine, mters 121°. \rightarrow glycyl-*dl*-*o*-tyrosine + 2 H₂O decomps. 120-30°, *o*-tyrosine + Me-CHCH-CHBrCOBr \rightarrow *dl*- α -bromosuccinyl-*dl*-*o*-tyrosine, mters 94°. \rightarrow *dl*-leucyl-*dl*-*o*-tyrosine. Similarly, *dl* *m*-tyrosine \rightarrow chloroacetyl-*dl* *m*-tyrosine, m 130°. \rightarrow glycyl-*dl* *m*-tyrosine, m 120-30° (foaming), *dl* *m*-tyrosine \rightarrow *dl*- α -bromosuccinyl-*dl* *m*-tyrosine, mters 127°. \rightarrow *dl*-leucyl-*dl* *m*-tyrosine, m 120-30° (foaming). 2-nitrotyrosine Et ester \rightarrow chloroacetyl-2-nitrotyrosine Et ester, mters 92°. \rightarrow chloroacetyl-2-nitrotyrosine, mters 165-6°. \rightarrow glycyl-2-nitrotyrosine, mters 220° (decompn). 1 2-nitrotyrosine ester \rightarrow *dl*- α -bromosuccinyl-2-nitrotyrosine ester, mters 76°. \rightarrow *dl*- α -bromosuccinyl-1-nitrotyrosine, mters 170°. \rightarrow *dl*-leucyl-1-nitrotyrosine, mters 210-20° (decompn). *dl* phenyl β -alanine \rightarrow chloroacetyl-*dl* phenyl β -alanine, m 141°. \rightarrow glycyl-*dl* phenyl β -alanine, m 222° (decompn). *dl* phenyl β -alanine \rightarrow *dl*- α -bromosuccinyl-*dl* phenyl β -alanine, m 132°. \rightarrow *dl*-leucyl-*dl* phenyl β -alanine, m 200° (decompn). All of the above dipeptides were resistant to Δ alkali at 37°, showing practically no cleavage in 120 hrs. Chloroacetyl-nitrotyrosine was readily hydrolyzed. Erepsin attacked only the 2 peptides of *p*-tyrosine. An erepsin-free prepn of trypsin kinase also attacked these 2 peptides, the relative extent of cleavage varying with the age of the enzyme prepn. A prepn 24 hrs old hydrolyzed leucyltyrosine more rapidly and glycyltyrosine less rapidly than the fresh prepn. The previously reported cleavage of chloroacetyl-*l*-tyrosine by trypsin kinase was confirmed, but the analogous bromosuccinyl derivative was found to be resistant, as were all the remaining haloacetyl derivatives and their corresponding dipeptides.

A. W. Diox

Preparation of dimethylacetacetic ester and of 2,2,2-dimethylbuten-1-ol. KARL FOLKERS AND HOMER ADKINS. *J Am Chem Soc* 53, 1416-9 (1931).—Details are given for the prepn of AcCHMeCO₂Et (70% yield) and of AcCMe₂CO₂Et (34% yield). Reduction of Et 2,2,2-dimethylbutyrate with Na and EtOH, using the method of Adams and Marvel (*C A* 14, 1677) gives 61.8% of 2,2,2-dimethylbuten-1-ol, b 126.5-31°. There was also recovered 7.9% of the Na salt of the acid. C. J. West

Rate of hydrogenation of acetoacetic ester, dehydroacetic acid, benzene, phenol and aniline over nickel at pressures from 27 to 350 atmospheres. HOMER ADKINS, HOWARD I. CRAMER AND KATHY CONNOR. *J Am Chem Soc* 53, 1402-5 (1931).—The relationship of pressure to the rate of catalysis by Ni of the hydrogenation of AcCH₂CO₂Et (I), dehydroacetic acid (II), C₆H₆, PhOH and PhNH₂, has been found to differ greatly among these 5 compds. It was entirely feasible to reduce I at pressures in the vicinity of 30 atm but the rate of reduction was greatly increased by increases in pressure, especially in the range from 120 to 370 atm. II reduced more than twice as rapidly at 149 as at 108 atm and about 4 times as rapidly at 323 as at 108 atm. It was found impractical to reduce PhNH₂ in the 30-atm range. Increasing the pressure greatly increased the rate of hydrogenation but increasing the pressure in the higher ranges up to 350 atm was not found to be particularly advantageous. The rate of reduction of PhOH and of C₆H₆ proceeded well in the 30-40 atm range, increased with pressure to the 150-170 atm range but was not sensitive to further increase in pressure up to 330 atm. C. J. West

Reactions of malonic esters with formaldehyde. H. KENNETH N. WALSH. *J Chem Soc* 1931, 653-7, cf *C A* 24, 2429.—The reaction of malonic esters with HCHO is obviously of the 2nd order, the initial speed of reaction, *r*, is approx proportional to the first power of the concn of each reactant separately. Within the limits that could be conveniently investigated, *k* is inversely proportional to the H ion concn

Piperidine has a profound effect on the reaction velocity, while Me_3N and NH_3 are without action within the limits of exptl error, the action of Me_3N appears to be temporary (possibly due to the destruction of the catalyst) C J WEST

Preparation of ethyl methylenedimalonate. KENNETH N WELCH *J Chem Soc* 1931, 673-4. cf *C A* 24, 2429— $\text{CH}_2(\text{CO}_2\text{Et})_2$ (36 g) and 3 g paraformaldehyde, heated on the water bath and treated with 10% EtOH KOH drop by drop until reaction commenced and then heated 6 hrs., gave 5 g $\text{CH}_2(\text{CO}_2\text{Et})_2$ and 28 g Et methylene dimalonate C J WEST

The structure and properties of crystal hydrates of oxalic acid. M A RAKUSIN *Chem.-Zig* 55, 128(1931)—R discusses the structure of oxalic acid from the point of view of his theory of 'hydroxyl loading' (cf *C A* 25, 69) He concludes that the dihydrate is not $\text{C}_2(\text{OH})_2$, as suggested by Mendeleev and by Genth, but is $(\text{CO}_2\text{H})_2 \cdot 2\text{H}_2\text{O}$ It has been incorrectly stated that cryst. oxalic acid begins to effloresce at 20° A sample on a watch glass did not lose a trace of H_2O in 29 days B S LEVINE

Chemistry of the glutamic acids. XXII. Optically active α,γ -dimethylglutamic acid. TERENCE H MCCORMS, JOHN PACKER AND JOCELYN F THORPE *J Chem Soc* 1931, 547-60 cf *C A* 21, 1257—*trans* $\text{HO}_2\text{CCH}(\text{Me})\text{CH}(\text{MeCO}_2\text{H})$ (I), m 147° , is easily resolved by strychnine in Me_2CO or Me_2CO and Et_2O strychnine *H I* ppts. on cooling, m $208-9^\circ$, then solidifies and m $20-3^\circ$ and then decomp. dil NH_4OH gives *I*, m $132.5-3^\circ$, $[\alpha]_D^{25} -100 \pm 1^\circ$ (H_2O , c 1 g in 100 cc), -106° (H_2O , c 4 g in 100 cc), $[\alpha]_D^{25} -136^\circ$ (EtOH, c 0.45 g in 100 cc) solns of *I* in H_2O and dil. HCl , NH_4OH and NaOH undergo no change in rotation at 25° for 3 days. The *d I* contained approx 25% *d I*, because of racemization in hot solns. of the strychnine salt. Brucine, quinine, quinidine, cinchonine and *d* MeCHPhNH_2 gave salts which either would not crystallize or were otherwise unsuitable for the resolution of the acid. Re-combination of the active forms gives the *d I* acid, m 147° The action of strychnine on *d I* in insufficient and sufficient boiling Me_2CO to dissolve the salts formed, in a cold $\text{Me}_2\text{CO}-\text{CHCl}_3$ mixt. and in sufficient boiling H_2O to dissolve the salts formed and of brucine in sufficient boiling Me_2CO to dissolve the salts formed is described. Racemization of *I* in solvents is reported as follows (solvent, temp., k (hr^{-1}) and half-change period are given) H_2O , 100° , 0.0388, 17.8 hrs., N HCl , 101.5° , 0.0323, 21.3 hrs., N HCl , 101° , 0.128, 5.4 hrs., N NaOH , 101.5° , 1.70, 24.5 min., Me_2CO , 56.5° , about 0.0088, about 80 hrs. Mutarotation or racemization of *I* also takes place more rapidly in hot solvents in the presence of strychnine than in its absence and even occurs very slowly in its presence in the cold, in boiling H_2O , k 0.577 hr^{-1} , half-change period 72 min.; in boiling acetone the values are 0.451 and 0.2. *I* and AcCl heated at 100° for 6 hrs., give only 7% of the hydroxy arylidide, m 210° . Strychnine *H glutamate*, m 195° (decompn.); fractionation of this salt showed no resolution. C J WEST

Structure of the glutamic acids and esters. L. Cyanoglutaric esters. GEORGE A. R. KOV AND HOMI R. NAYR *J Chem Soc* 1931, 560-75—The cyanoglutaric esters display tautomerism of the same type as that of many other unsatd. cyano esters. They can in some cases be isolated in 2 distinct forms, the reactions of these isomers are fully accounted for by the α,β - and β,γ -unsatd. formulas and the "normal" formulation suggested for some of them is no longer necessary. These isomers are interconvertible in the presence of suitable catalysts, forming equil. mixts., notably when they are liberated from their Na derivs. and for this reason the method of distinguishing between "normal" and "labile" isomers used by Thorpe and Bland is not effective. This statement applies only to α -cyano- and α -carboxyglutaric esters, because the terms "normal" and "labile" have also been used to denote the esters of the glutamic acids themselves. There are no doubt distinct and their isomerism presents a separate problem in which stereoisomerism unquestionably plays an important part. $\text{EtO}_2\text{CCH}(\text{CN})\text{CMe}(\text{CHCO}_2\text{Et})$ [equil. mixt. with $\text{EtO}_2\text{CC}(\text{CN})\text{CMe}(\text{CH}_2\text{CO}_2\text{Et})$ (I) was prep'd. by the following methods. Rogerson and Thorpe's method (*J. Chem. Soc.* 87, 1669), 33% excess of EtOH removed *in vacuo* at 60° , 35%, EtONa prep'd. from "mol." Na in dry Et_2O with theoretical quantity of EtOH, 35%, K in place of Na, 56%, 20 g K in 150 cc. EtOH, cooled, is treated with 56 g $\text{NCCCH}_2\text{CO}_2\text{Et}$, then with 65 g $\text{AcCH}_2\text{CO}_2\text{Et}$ and kept at room temp. for 2 days, 80%; the yield with triethanolamine or piperidine is only 10 and 3%, resp. *I* b_p 170° , d_4^{25} 1.0897, n_D^{25} 1.45917. Oxidation of *I* with O_3 gives OHCCO_2Et , $\text{AcCH}_2\text{CO}_2\text{Et}$, $\text{AcCH}(\text{CN})\text{CO}_2\text{Et}$ and $\text{OC}(\text{CN})\text{CO}_2\text{Et}$, whose phenylhydrazones, yellow, m 127° . The pure β,γ -ester could not be prep'd. The pure α,β -ester, *Et* α -cyano- β -methyl- Δ^2 -propene- α,γ -dicarboxylate (II), b_p $178-80^\circ$, d_4^{25} 1.0918, n_D^{25} 1.47409, *I* addn. 3% in 1 hr., results from the K salt of *I* with BrOH or HCl or by the action of EtOH EtOK on *I*; O_3 gives a mixt. of $\text{AcCH}_2\text{CO}_2\text{Et}$ and OC -

(CN)C₂H₅Et. EtONa gives the equal ester, the equal mixt. contains about 23% of a β ester. The K deriv. of II and MeI give the β , γ -di Me deriv. of II, bp 162°, d₄²⁰ 1.146, n_D²⁰ 1.49614, oxidation gives AcCHMeCO₂Et and OC(CN)CO₂Et. EtO-C(CN)C(CN)CMe(CN)MeCO₂Et (equal ester), bp 169°, d₄²⁰ 1.0708, n_D²⁰ 1.46095, γ gives AcCHMeCO₂Et + MeCOCO₂Et and AcCH(CN)CO₂Et. The pure β , γ -compd. 1.17/CCH(CN)CMeCMeCO₂Et results by the action of BrOH on the K deriv. or with 1.16K and bp 154° d₄²⁰ 1.0729, n_D²⁰ 1.46915, the structure was established by oxidation. The γ benzyl deriv., bp 214°, d₄²⁰ 1.0997, n_D²⁰ 1.51890, γ adds O, the action of 1.16Na gives the normal and labile esters as described by Thorpe and Wland. Et α -cyano α -methyl β propene α , γ dicarboxylate, bp 160-1°, d₄²⁰ 1.0602, n_D²⁰ 1.45415 does not react with ICl (70% yield), O₂ gives OHCCO₂Et. The α -Et deriv. bp 158-60°, d₄²⁰ 1.0603, n_D²⁰ 1.5324 (25% yield) EtOCCN(CN)CHCMeCO₂Et bp 166°, d₄²⁰ 1.0743, n_D²⁰ 1.47933, I adds 89% in 1 hr.; the use of EtOK in the cold gives a 65% yield, the color produced by aq. FeCl₃ quickly fades; oxidation shows the presence of the α δ and β esters. Et α -cyano α , γ -dimethyl β propene α , γ dicarboxylate bp 176°, d₄²⁰ 1.0729, n_D²⁰ 1.49818 (70% yield), oxidation gives AcCO₂Et and a phenylhydrazine bull. m 237-8°, doubtless from the ester OHCCMe(CN)C(CN)Et. LiONa gives Li₂CO₃ and probably Et α -cyano α , γ -dimethylacrylamide, bp 169°.

The asymmetry of *meso*-tartaric acid. J. ROSENBERG. Pharm. Weekbl. 69, 246-8 (1931).—A theoretical discussion of the possible asymmetry of *meso*-tartaric acid and certain objections to the hypothesis proposed by Scheninger [C. A. 25, 1893].

W. D. DOX
Theoretical and experimental study of a high-yield technical urea process. A. KALLERNE. Z. Elektrochem. 35, 799-802 (1930).—Urea may be prepd. by a method represented by $2\text{NH}_3 + \text{CO}_2 \rightleftharpoons \text{CO}(\text{NH}_2)_2 + \text{H}_2\text{O}$ (I) between a temp. range of 100° to 130°. At equal the urea content calcd. on carbamate is 33-43%. The equil. is affected decidedly by the H₂O so as to prohibit greater yields. Another method which has been previously studied from the standpoint of constitution of NH₂ thiocarbamates is represented by the equation $\text{COS} + 2\text{NH}_3 \rightleftharpoons \text{CO}(\text{SNH}_2)_2 + \text{H}_2\text{S}$ or $\text{COS} + 2\text{NH}_3 \rightleftharpoons \text{CO}(\text{NH}_2)_2 + \text{H}_2\text{S}$ (II). K. studied the kinetics and quant. aspects of these methods, concn. more particularly upon the 2nd. The free energy of the reactions was calcd. from the knowledge of the free energies of the individual reactants. The pressures in atm. with which urea is formed in both cases are as follows: I, 1.9×10^{-4} at 25°, 2.1×10^{-3} at 77°, 0.57 at 100° and 0.74 at 130°; II, 42 at 25°, 4.10^4 at 100° and 3.10^4 at 130°. The enormous differences in pressures are not sufficiently accounted for by the differences in the reaction velocities. A more complete account will appear in Z. anorg. allgem. Chem.

W. VANSLOW
Structure of the urea molecule. L. ECKART. Ber. 64B, 679-81 (1931), cf. C. A. 17, 1579.—According to Devoto (C. A. 25, 626) urea and thiourea in H₂O have the structure represented by the "zwitter-ion" formulas $\text{NH}_2^+\text{C}(\text{NH}_2)\text{NH}_2^-$ and $\text{NH}_2^+\text{C}(\text{NH}_2)\text{NH}_2^-$. Arndt (C. A. 25, 2145) suggests another polar formula, $\text{NH}_2^+\text{C}(\text{NH}_2)\text{S}^-$. E. asserts that there is no basis for the assumption of a polar formula for urea because: (1) analysis of the dielec. polarization of aq. solns. gives no proof of the existence of an especially strongly polar form of urea. (2) "zwitter ion" formation can occur only if the product of the 2 dissociation constants of the neutral form of urea is greater than 10^{-14} , this requirement can be fulfilled neither by the formula $\text{CO}(\text{NH}_2)_2$, nor by the formula $\text{H}_2\text{N}(\text{C}(\text{NH}_2)\text{NH}_2)\text{OH}$. (3) the crystal mol. of urea is very sym. and there is no evidence that dissolving urea in H₂O causes noticeable change in its structure.

LOUISA KELLER
Configuration of the doubly-linked tervalent nitrogen atom. IV. The resolution of β -methylmethacrylate dithiocarbonate carboxyphenylhydrazones. WM. H. MALLS and BENJAMIN C. SATCHEL. J. Chem. Soc. 1931, 837-46.—Mills and Schindler (C. A. 17, 1628) resolved the pyridylhydrazones of cyclohexene dithiocarbonate and concluded that its mol. dissymmetry was to be referred to the non planar distribution of the valencies of the doubly linked N atom. Because of the absence of definite proof as to whether this has the *cis* or the *trans*-configuration, work was undertaken to prep. a doubly linked tervalent N compd. which would be better fitted to provide definite evidence of the non planar distribution of the valencies of the N atom in this state of conformation. EtONa, satd. with H₂S, C₂S added and the soln. boiled for 0.5 hr. and then treated with MeCH(CH₃)Et₂ and allowed to stand 24 hrs., gives 50% of β -methylmethacrylate trithiocarbonate (II), yellow, m 74°. I and EtOH-PhNHNH₂,

With the Na compd of $\text{CH}_2(\text{CO}_2\text{Et})_2$, there results Et cyclopentane 1,1-dicarboxylate; $[\text{NaC}(\text{CO}_2\text{H})_2]_2$ gave a mixt of *cis*- and *trans* hexahydrophthalic acids when the tetra- CO_2H acid was converted by heating to the di- CO_2H acid $\text{CH}_2[\text{NaC}(\text{CO}_2\text{Et})_2]_2$ gives a mixt of $(\text{CH}_2)_4\text{C}(\text{CO}_2\text{Et})_2$ and a polymer of $\text{CH}_2\text{C}(\text{CO}_2\text{Et})_2$ (Zelinskii, *Ber.* 22, 3296(1889))

C. J. WEST

1-Benzylcyclohexene. Condensation of benzyl alcohol with cyclohexene under the action of phosphoric anhydride. D. N. KURSAKOV *J. Russ Phys.-Chem. Soc.* 62, 1691-5(1930).—1 Benzylcyclohexene (I) was obtained by dehydration with P_2O_5 of PhCH_2OH in cyclohexene $\text{C}_6\text{H}_{10} + \text{HOCH}_2\text{Ph} \rightarrow \text{C}_6\text{H}_9\text{CH}_2\text{Ph}$. Another product of the reaction is a resinous mass which cannot be distd. without decompn. To 80 g cyclohexene and 50 g P_2O_5 cooled with ice 30 g PhCH_2OH was added by drops. The mixt was, after 12 hrs' standing, treated with NaHCO_3 soln., washed with water, dried with Na_2SO_4 and fractionated, the collected fraction b₁ 127-2-84°, d₄ 0.8634, n_D 1.5400. Other const. as well as analysis show the product to be identical with I as described by its earlier investigators. K. assumes that the reaction takes place in 2 steps: (a) benzyl phosphate combines with cyclohexene at the double bond, (b) H_3PO_4 is liberated. Partial oxidation of I by means of alk. KMnO_4 at 0-4° results in formation of what K. believes to be the glycol $\text{C}_6\text{H}_9(\text{OH})_2\text{CH}_2\text{Ph}$ and BrOH .

J. C. TOLPIN

Preparation of highly active aluminum chloride. K. BODENDORF *J. prakt. Chem.* 129, 340(1931).—Five g. Al powder, 1 g. HgCl_2 and 0.5 g. I in 100 cc. C_6H_6 are warmed nearly to boiling and dry HCl passed in with stirring, the H_2O both may be removed as soon as H evolution starts, the Al is not completely consumed since it is covered with AlCl_3 , the product is ready to use when the H evolution ceases.

C. J. WEST

Reaction of aliphatic aldehydes according to Friedel-Crafts. K. BODENDORF, *J. prakt. Chem.* 129, 337-9(1931).—Anhyd. AlCl_3 (12 mols.) in 300 cc. C_6H_6 is satd. with HCl at 0° and then treated slowly with 1 mol. aldehyde in 100 cc. C_6H_6 , during 2 hrs. and then stirred 1-2 hrs. The C_6H_6 free residue (150 g.) from AcH gave 8 g. PhEt , 88 g. MeCHPh , and 33 g. of a yellow oil from which was recovered 4.5 g. of 9,10-dimethyl-9,10-dihydroanthracene (I). MeCHPh , (10 g.) and 20 g. AlCl_3 in 100 cc. C_6H_6 heated 1 hr. give 1.5 g. PhEt and 4 g. I, 10 g. of the oil, heated with AlCl_3 , gives 1 g. MeCHPh and 4 g. I. The residue from EtCHO (154 g.) gave 16 g. PhPr , 91 g. EtCHPh , and 36 g. high boiling products. The residue from PrCHO (173 g.) gave 29 g. PhBu , 82 g. PrCHPh , and 33 g. yellow oil, b₁ 175-240°. The residue from iso-PrCHO (162 g.) gave 21 g. iso-BuPh , 102 g. Me_2CHCHPh , and 29 g. yellow oil, b₁ 170-240°, 20 g. of the latter with AlCl_3 heated 1 hr., gave 7 g. iso-BuPh and 5 g. Me_2CHCHPh .

C. J. WEST

Electric moments of some organic molecules in benzene solution. A. C. EIDE AND O. HASSEL, *Tids. Kjemiog. Bergesen* 10, 93 5(1930), cf. following abstract.—A series of results of detns. of elec. moments of various org. molts. in C_6H_6 soln. is given. The small value obtained for $p\text{-O}_2\text{NC}_6\text{H}_4\text{CN}$ (0.72×10^{-18}) seems to be due to a neg. effect of the N when assocd. with C. The 2 hydrobenzoic acids were investigated as examples of 2 compds. standing to each other in the relationship of *meso*- and *dl* form. The moments are almost the same. The dipole moment of *p*-chloroaniline is 2.9×10^{-18} while the calcd. value is about 3.0×10^{-18} .

B. C. A.

Electric moments of organic molecules. IV. O. HASSEL AND E. NAESHAGEN, *Tids. Kjemiog. Bergesen* 10, 81-4(1930), cf. C. A. 24, 2346-7.—The structure of ring systems is discussed on the basis of the elec. moments of the molts. Results are given for a no. of cyclic compds. in C_6H_6 soln., and these are in general agreement with those of other observers. The moments of some methylated sugars are given, but from these data it is not possible to arrive at any accurate conclusions concerning the structure of sugars. For thiophene, the max. value found for the elec. moment was 0.63×10^{-18} , which may be compared with that for furan. The moment of carbene is 3.17×10^{-18} , which is somewhat different from that for methene (2.80×10^{-18}).

B. C. A.

Electric moments of organic molecules. VI. Ortho-effect of derivatives of *p*-dichlorobenzene and of the *sym*-trichloro- and tribromobenzenes. O. HASSEL AND E. NAESHAGEN, *Z. physik. Chem.* Abt. B, 12, 79-88(1931), cf. preceding abstr.—The position of substituents in benzene rings and in C_6H_5 may be detd. by measuring the dipole moment in benzene soln. The dipole moments were detd. for 2,5- $\text{Cl}_2\text{C}_6\text{H}_3\text{NO}_2$, *o*-, *m*-, *p*-dichloroaniline, 2,4,6-trichloro- and tribromoaniline, *o*- $\text{Cl}_2\text{C}_6\text{H}_3\text{CN}$, 2,5- $\text{Cl}_2\text{C}_6\text{H}_3\text{CN}$, 2,4,6- $\text{Cl}_3\text{C}_6\text{H}_2\text{CN}$, 2,4,6- $\text{Cl}_3\text{C}_6\text{H}_2\text{OH}$, 2,4,6- $\text{Br}_3\text{C}_6\text{H}_2\text{OH}$, 1,2,4- $\text{C}_6\text{H}_3\text{Cl}_3$, 1,2,4,6- $\text{C}_6\text{H}_2\text{Cl}_4$ and $\text{C}_6\text{H}_2\text{Br}_4$, 2,4- $(\text{O}_2\text{N})_2\text{C}_6\text{H}_3\text{Cl}$. Derivs. of *sym*- $\text{C}_6\text{H}_2\text{Cl}_4$ and $\text{C}_6\text{H}_2\text{Br}_4$ showed the ortho-effect distinctly.

FRANK URBAN

Nitridation studies. III. Phenyl iodide dichloride and the *N*-chloroacid amides as nitridizing agents. A LAURENCE CURL AND W. CONRAD FERNELIUS *J. Am. Chem. Soc.* 53, 1478–82(1931) cf *C A* 24, 3498— PhICl_2 and ICl may be used as nitridizing agents in liquid NH_3 , but the *N*-Cl acid amides so far investigated are valueless in this connection. PhICl_2 reacts with NH_3 (liquid or gas) to give N , PhI and NH_4Cl . The *N*-Cl acid amides react with NH_3 to give the parent amides, NH_4Cl , N and, under certain conditions, tarry decomposition products. In CHCl_3 , PhICl_2 , dichloramine-*T*, AcNClPh and *N*-chlorosuccinimide convert $(\text{PhNH})_2$ into a mixt. of $(\text{PhN})_2$ and $(\text{H}_2\text{NC}_6\text{H}_5)_2$, 2 HCl .

C. J. WEST

1,3,5-Trinitrobenzene or benzoate. LOUIS DESVERGNES *Chimie & Industrie* 25, 3–17, 291–306(1931)—The properties and prepn of 2,4,6-(O_2N) $_3\text{C}_6\text{H}_2\text{Cl}$ and of 1,3,5- $\text{C}_6\text{H}_2(\text{NO}_2)_3$ are reviewed. A bibliography of 80 references is given. Expts undertaken to det. the most suitable process of manu. of these 2 compds are described. Comparison of the Griesheim Chemische Fabrik process (Ger. pat. 78 309), Desvergnès process (modification of the preceding process in which the amt. of HNO_3 is considerably reduced) and Frankland and Garner process (*C A* 14, 3068) for the prepn. of picryl chloride showed that, from the standpoint of cost and consumption of raw materials, the Desvergnès and Frankland-Garner processes yield a technically pure product at practically the same cost, which is much lower than that of the Griesheim process. Comparison of the prepn. of 3 $\text{C}_6\text{H}_2(\text{NO}_2)_3$ by (a) reduction of picryl chloride, (b) oxidation of trinitrotoluene with $\text{K}_2\text{Cr}_2\text{O}_7$ and (c) nitration of *m*- $\text{C}_6\text{H}_4(\text{NO}_2)_2$ showed that (b) is the most advantageous from the doublest standpoint of the purity of the product and cost of production.

A. PAFINEAU COUTURE

Determination of constitution by changing the order in which groups are introduced into the benzene nucleus. J. W. DIENSKER *Rec. trav. chim.* 50, 21–31(1931), cf *C A* 25, 922—4- $\text{NCSC}_6\text{H}_4\text{NH}_2$ was converted by means of the Sandmeyer reaction into 4-chlorothiocyano benzene, *m* 35–6°, which was obtained also from *p*- $\text{ClC}_6\text{H}_4\text{NH}_2$ by diazotization and treatment with CuSCN . $\text{Co}(\text{SCN})_3$ gives better results in this reaction than CuSCN , but much better results are always obtained in reactions of this kind by the method of Thurnauer (*Ber.* 23, 760(1890)) namely, by taking a soln. of CuSCN in a concd soln. of KCNS . Application of the Sandmeyer reaction to 4- $\text{NCSC}_6\text{H}_4\text{NH}_2$ also gave 4-bromothiocyano benzene *m* 54° (Challenger and Collins, *C A* 18, 2689 give 50°), which could be prepd. also from diazotized *p*- $\text{BrC}_6\text{H}_4\text{NH}_2$ by Thurnauer's method. 4-Iodothiocyano benzene, *m* 51°, was prepd. from 4- $\text{NCSC}_6\text{H}_4\text{NH}_2$ in the usual way, but could not be obtained from 4- $\text{IC}_6\text{H}_4\text{NH}_2$ on treating the diazotized soln. with CuCNS , which is in accordance with the experience of Challenger and Collins (*C A* 18, 2689). Detailed directions are given for the prepn. of 4-cyanothiocyano benzene, *m* 127.5°, and of 1,4-dithiocyano benzene, *m* 108.5°, from 4- $\text{NCSC}_6\text{H}_4\text{NH}_2$, on nitration with HNO_3 and H_2SO_4 , the latter compd. gives 2-nitro-1,4-dithiocyano benzene, *m* 146.5° (cf. Challenger and Peters, *C A* 22, 3152) which was now obtained also by means of abs. HNO_3 . The hydrolysis of 4- $\text{NCSC}_6\text{H}_4\text{Br}$ with boiling 10% KOH gave (*p*- BrC_6H_4) $_2\text{S}$, *m* 93°, together with 4- $\text{BrC}_6\text{H}_4\text{SH}$, the same compds being formed on boiling with 75% H_2SO_4 . The compd. remained unchanged, however, on boiling with 10% H_2SO_4 . On hydrolyzing 4- $\text{NCSC}_6\text{H}_4\text{I}$ with boiling 10% NaOH , (*p*- IC_6H_4) $_2\text{S}$, *m* 125°, and 4- $\text{IC}_6\text{H}_4\text{SH}$, *m* 82–6°, were obtained. These reactions are in accordance with the general rule that NCS compds. are decompd. by heating with acids or alkalis with the formation of mercaptans and disulfides (cf. Brüning, *Ann.* 104, 198(1857)).

C. F. VAN DUTY

The nitration of 4-thiocyanacetanilide. J. W. DIENSKER *Rec. trav. chim.* 50, 165–83(1931)—The nitration of 4- $\text{NCSC}_6\text{H}_4\text{NHAc}$ (I) was studied in order to investigate whether the NCS group is displaced in nitration processes like the SO_3H group (cf. Zincke, *Ann.* 339, 202(1903), de Lange, *C A* 20, 1982). Challenger and Peters (*C A* 22, 3152) nitrated this compd. with HNO_3 and H_2SO_4 to 2,4- $\text{O}_2\text{N}(\text{NCS})\text{C}_6\text{H}_2\text{NHAc}$, D, using abs. HNO_3 , obtained 2,6-dinitro-4-thiocyanophenyl nitramine (II), the NCS group not being displaced. At first an attempt was made to det. the constitution of this compd. by converting 4-thiocyananisole or -phenetole into the 2,6-dinitro deriv., which was further transformed into the amino compd. The latter, however, could not be transformed into the nitramine. The NCS group was then oxidized to a SO_3H group, which was subsequently displaced by a nitro group, finally the 2,4,6-(O_2N) $_3\text{C}_6\text{H}_2\text{NHNO}_2$ being hydrolyzed to picric acid. I was nitrated with abs. HNO_3 at -10° to II, decomps. explosively at 70° . II is rather unstable and was converted into a sticky mass on keeping in a bottle for 4 weeks. 4-Thiocyanophenetole (III), *m* 47.5–8°, was obtained from *p*-phenetidine in the usual way together with some of the isomeric mustard oil to be described later, the ethylation of 4- $\text{NCSC}_6\text{H}_4\text{OH}$

(Söderback *C A* 14, 1808), however, did not give III, but an unknown substance which could not be identified. On nitrating III with 65% HNO_3 , 2-nitro-4-thiocyanophenol, m 85° (IV), is formed, while from the mother liquors of the recrystn of IV, some 4-nitrophenol, m 58°, was obtained. A 2nd by product, formed in this nitration, is picric acid, formed by further nitration and hydrolysis of the 4-nitrophenol. The constitution of IV was proved by its formation from 2,4- $\text{O}_2\text{N}(\text{H},\text{N})\text{C}_6\text{H}_3\text{OEt}$ (Reverdin, *C A* 21, 1451) by the Sandmeyer reaction. Nitration of IV with abs HNO_3 at 0° gives 4-thiocyano-2,6-dinitrophenol (V), m 77°, the constitution of which was established by converting it into picric acid, first the SCN group was oxidized to the SO_3H group with HNO_3 , then the OEt group was hydrolyzed with alkali to the OH group and the resulting 4,3,5- $\text{HO}(\text{O}_2\text{N})_2\text{C}_6\text{H}_2\text{SO}_3\text{H}$ nitrated to picric acid. On warming with alc. NH_3 , V is converted into 4-thiocyano-2,6-dinitroaniline (VI), m 180°. On carrying out the Sandmeyer reaction with *p*-anisidine, 4-thiocyanoanisole (VII), m 35°, is obtained, together with some of the isomeric mustard oil, to be described later. In the nitration of VII, 2-nitro-4-thiocyanoanisole was formed together with some 4-nitroanisole and some picric acid, while further nitration affords 4-thiocyano-2,6-dinitroanisole, m 93°, which was again converted into VI by means of alc. NH_3 . VI was recovered unchanged after treatment with abs HNO_3 , while II was not converted into VI on heating with phenol by van Romburgh's method, a substance, m 200-1°, which was not investigated further, being obtained instead of VI. The constitution of II was proved, however, by boiling II with HNO_3 , hydrolyzing the SO_3H acid with alkali to the corresponding phenol and nitrating the latter to picric acid. Finally, a discussion of the m ps of the various thiocyan compounds in connection with the m p rules postulated by Franchimont (*Rec trav chim* 16, 126(1897)), is given.

C F VAN DUIN

Hydrogenation of aniline under pressure in presence of activators (cerium and lanthanum). V S SADIKOV AND A YU SHAGALOV. *J. Russ Phys.-Chem. Soc.* 62, 1635-42(1930).—In search for a new and cheaper catalyst, a series of expts on hydrogenation of aniline under pressure were carried out by S and S, using as catalysts Os and Ni with addn of activating agents (Ce and La). The results for Os tend to show that addn of Ce has no influence on the duration of hydrogenation, decreases the quantity of benzene 3 times of cyclohexylamine twice, increases twice the quantity of dicyclohexylamine, and by 10% that of non-reacting aniline. The quantity of H consumed is somewhat smaller. With Ni + Ce the addn of Ce renders the catalyst more sensitive to thermal decompn on prolonged heating in the process of its reduction. Although the system Ni + Ce cannot substitute Os it is a more efficient catalyst than Ni. For Ni + La a short time of preheating is more important than for Ni. La does not possess the condensing properties of Ce and the reaction results in formation of more cyclohexylamine (70.21% compared to 33%) and less dicyclohexylamine (17% against 52%). In general the greater efficiency of Os is based on its ability for reduction at room temp as heating is shown to poison Ce and La.

J G TOLPIN

Interaction of alcohols and ethers with aniline hydrochloride. S A BUSE AND A I TRAVIN. *J. Russ Phys.-Chem. Soc.* 62, 1685-90(1930).—It has been shown many times that chlorozincaniline with aliphatic alcs gives at high temps and pressures mostly homologs of aniline substituted in the ring. Halogen salts of aniline or its homologs give secondary and tertiary bases. B and T worked out a method of obtaining benzylaniline and octylaniline and investigated the interaction of benzyl and octyl alc with aniline-HCl and obtained BzH and $\text{C}_8\text{H}_{17}\text{ClHO}$ by oxidation of the corresponding anilines. Heating at low temp for a short time gave good results. It is important to drive off the water as it is formed. PhCH_2OAc gives a higher yield than PhCH_2OH . For octylaniline the reverse is true. Benzylaniline was obtained by heating in an oil bath 10 g benzyl alc with 20 g PhNH_2HCl 4-5 hrs at 180-90°, cooling, treating with NaOH, washing, drying with fused Na_2SO_4 and fractionating. Yield, 59%. By using 20 g PhCH_2OAc and 35 g PhNH_2HCl the yield is 79.3%. Oxidation of benzylaniline to BzH is done most effectively by ferric salts and $\text{ONC}_4\text{H}_9\text{NMe}_2$. To 20 g PhNHCH_2Ph 150 g of $\text{Fe}_2(\text{SO}_4)_3$ was added in 10 g portions and in presence of H_2SO_4 . BzH was distd, extd with ether and treated with a satd soln of bisulphite. Yield 61.7%. FeCl_3 gave only a 51.7% yield. Oxidation by means of $\text{ONC}_4\text{H}_9\text{NMe}_2$ is effected by adding 8 g of it to 5 g of benzylaniline and slowly heating the mixt. to 150° for 1 hr, using an air condenser. An excess of dil H_2SO_4 is added and BzH distd. Yield, 73.8%. $\text{ONC}_4\text{H}_9\text{NMe}_2$ is reduced to azoxydimethylaniline ($\text{PhCH}_2\text{NHPH} + 2\text{ONC}_4\text{H}_9\text{NMe}_2 = \text{PhCHNPH} + (\text{Me}_2\text{NC}_4\text{H}_9)_2\text{N}_2\text{O} + \text{H}_2\text{O}$), the latter, m 241°, was isolated from the reaction mixt. by extg the sol part with alc, washing and recrystg from benzene and CHCl_3 . Octylaniline was prepd

by heating a mixt of 50 g octyl alc and 100 g aniline-HCl on an oil bath 7-8 hrs to 210-40°, treating with NaOH, washing with water, drying with Na_2SO_4 and rectifying. Yield, 75%, b_p 146-8°. $\text{C}_{11}\text{H}_{13}\text{OAc}$ was also used instead of octyl alc at 270-80°. The product b_p 158-85° corresponds to the formula $\text{C}_{11}\text{H}_{13}\text{N}$, gives no cryst. salts and no isonitrile or diazo reaction. Oxidation with $\text{Fe}_2(\text{SO}_4)_3$ and $\text{ONC}_6\text{H}_5\text{NMe}_2$ gives octanal, the semicarbazone of which m 100°, it is therefore concluded that the compd is $\text{Me}(\text{CH}_2)_8\text{CH}_2\text{NHPh}$, d_{40}^{25} 0.9089, n_D^{25} 1.5132. J G TOLPIN

Reactions of some carbonyl compounds with phenylhydrazine. L CHAS. RAIFORD AND WILBUR T DADDOW *J Am Chem Soc* 53, 1552-8 (1931) —The CO radical in HCONHPh does not condense with PhNHNH_2 to give a hydrazone under the conditions studied. The products are PhNH_2 and HCHNHNHPh . Substitution products of HCONHPh react similarly. $(\text{PhNH})_2\text{CO}$ reacts with PhNHNH_2 to give a semicarbazide and eliminates PhNH_2 , the corresponding thiourea behaves in the same way *sym* $\text{Di}(2,5\text{-dibromophenyl})\text{thiourea}$ m 191° (Dyson, George and Hunter, *C A* 21, 1637, give 154°), the 2,4-di Br isomer, m 203°. The following *thiosemicarbazides* were prepd by heating the required urca with PhNHNH_2 in EtOH for 6 hrs: 1-phenyl-4-(4-bromophenyl), m 179-80°, 1-phenyl-4-(2,5-dibromophenyl), m 188°, 2,4-di-Br isomer, m 177-8°, 1-phenyl-4-(o-tolyl), m 170 1°, m tolyl isomer, m 173-4°, 1-phenyl-4-(a-naphthyl), m 192 3°, 1-phenyl-4-(3-nitrophenyl), m 172°. Oxidation of thioureas with PbO_2 in C_6H_6 gives carbodimides: *di* m tolyl, m 118-9°, *di*(2-bromophenyl), m 98-100°, *di* 4-bromophenyl, m 144°, *di* 2,4-dibromophenyl, m 156-7°, *di* 2,5-dibromophenyl, m 172 3°. With PhNHNH_2 in C_6H_6 these compds yield anilguanidines, *di* o tolyl, m 137°, *di* m tolyl, m 158 9°, *di* 2-bromophenyl, m 147-8°, *di* 3-bromophenyl, m 162-3°, *di* 4-bromophenyl, m 202-3°, *di* a naphthyl, m 157-9°, *di* β naphthyl, m 181-2°. Oxidation of these gives *azo* compds: *di*-o tolyl, pale red, m 113-4°, *di* m tolyl, brilliant red, m 107-8°, *di* 2-bromophenyl, brilliant red, m 132-3°, 3-Br isomer, dull red, m 131 2°, 4-Br isomer, red, m 163-4°. The evidence indicates that these guanidine derivs are not hydrazones. C J W

Reduction of azobenzene, azoxybenzene and nitrosobenzene by the system magnesium + magnesium iodide. W E BACIMANN *J Am Chem Soc* 53, 1524-31 (1931) — $(\text{PhN})_2$ is reduced by the system $\text{Mg} + \text{MgI}_2$ to $[\text{Ph}(\text{MgI})\text{N}]_2$. $\text{PhNN}(\text{O})\text{Ph}$ is reduced by the binary system to $(\text{PhN})_2$ or to $(\text{PhNH})_2$, according to the amt. of reducing agent employed. Substituted azobenzenes and azoxybenzenes react in the manner of the unsubstituted compds except that in certain instances reduction proceeds to a considerable extent beyond the hydrazobenzene stage. PhNO is reduced principally to $[\text{Ph}(\text{MgI})\text{N}]_2$, $(\text{PhN})_2$ is formed as an intermediate product in this reaction. The reactions of $[\text{Ph}(\text{MgI})\text{N}]_2$ with I, O, CO_2 , $\text{PhNN}(\text{O})\text{Ph}$, $(\text{PhCO})_2$ and BzH are described. The results obtained, in agreement with those of previous studies, lend support to the hypothesis that the active reducing agent is MgX , generated by interaction of Mg and Mg halide. C. J. West

Action of amines on 2-substituted semicarbazones. JAMES CHAPMAN AND FORSYTH J. WILSON *J Chem Soc* 1931, 507-14 — $\text{Me}_2\text{C NNPhCONH}_2$ (I), heated with PhCH_2NH_2 until the evolution of NH_3 had ceased (1 hr at 160°), gave a mixt. of $(\text{PhCH}_2\text{NH})_2\text{CO}$, m 171°, and $\text{PhCH}_2\text{NHCONH}_2$, m 149°. With PhNH_2 40 min at 155-60° it gives $(\text{PhNH})_2\text{CO}$ and a small quantity of a compd m 133°. $p\text{-MeC}_6\text{H}_4\text{NH}_2$, heated 40 min at 145-50°, gives $(p\text{-MeC}_6\text{H}_4\text{NH})_2\text{CO}$ and a small quantity of a compd m 162°. $\text{PhCH}_2\text{CH}_2\text{NH}_2$, at 150-60° gives $(\text{PhCH}_2\text{CH}_2\text{NH})_2\text{CO}$ and a tar. $\text{C}_6\text{H}_5\text{NH}_2$ 45 min at 160° gives $(\text{C}_6\text{H}_5\text{NH})_2\text{CO}$, m 91°, and a tar. Heating I at 140-5° for 1 hr. gives cyanuric acid and $\text{Me}_2\text{C NNHPh}$. Thus the reaction with the amines may be formulated as a thermal decompn., followed by the reaction of the $\text{Me}_2\text{C N NHPh}$ with the amine to give the $(\text{RNH})_2\text{CO}$ and NH_3 , hence none of the expected 2,4-disubstituted semicarbazone would be expected. PhMeC NNPhCONH_2 , m 126°, behaves similarly, PhNH_2 giving $(\text{PhNH})_2\text{CO}$ and PhMeC NNHPh . $\text{PhCH}_2\text{N}(\text{CH}_2\text{Ph})\text{CONH}_2$ is stable even after 15 hrs heating at 200°, since the reaction with PhNH_2 and $p\text{-MeC}_6\text{H}_4\text{NH}_2$ occurs at 165-70°, the products are $\text{PhCH NN}(\text{CH}_2\text{Ph})\text{CONHPh}$, m 152°, and benzaldehyde 4-*p*-tolyl 2-benzylsemicarbazone, m 172°, PhCH_2NH_2 at 175° gives some $(\text{PhCH}_2\text{NH})_2\text{CO}$ and benzaldehyde 2,4-dibenzylsemicarbazone, m. 124°. PhCH NNMeCONH_2 and PhNH_2 at 180° give PhCH NNMeCONHPh , m. 108°; apparently the decompn temp. of the semicarbazone is above 180°. With PhCH_2NH_2 at 165-70° for 1 hr the main product was $(\text{PhCH}_2\text{NH})_2\text{CO}$, only a small quantity of benzaldehyde 4-benzyl-2-methylsemicarbazone (II), m 106°, being formed, there also results a minute quantity of a compd, $\text{C}_{11}\text{H}_{15}\text{O}_2\text{N}_4$, m 88-9°. Heating for 3 hrs gives a small quantity of a compd, $\text{C}_{11}\text{H}_{15}\text{O}_2\text{N}_4$, m 63°. PhCH NNPhCONHMe

did not react with IhNH_2 at 180° . II and PhCH_2NH_2 at 170° give $(\text{PhCH}_2\text{NH})_2\text{CO}$. The mechanism of these reactions is discussed. C. J. WEST

Reactions of nitrosulfonyl chlorides. **II. Separation of nitrosulfonyl chlorides by means of hydrazine hydrate.** WM. DAVIS, I. R. DRAICK, R. STORRIE, AND STANLEY H. TUCKER. *J. Chem. Soc.* 1931, 624-9, cf. *C. A.* 23, 3661— $\text{O}_2\text{N}(\text{C}_6\text{H}_4)_2\text{SO}_2\text{NHNH}_2$ and MeCO give *acetone o-nitrobenzenesulfonylhydrazide*, m $147-8^\circ$ (decompn), *piperonal deriv*, bright yellow, m $177-9^\circ$ (decompn), the preps were carried out by adding the cold solid to the aldehyde or ketone and boiling 1 min. $\text{p-O}_2\text{N}(\text{C}_6\text{H}_4)_2\text{SO}_2\text{Cl}$ (4.43 g) in C_6H_6 , gradually added to $\text{N}_2\text{H}_4 \cdot \text{H}_2\text{O}$ in abs. EtOH at 40° , 0.5 N AcONa added and the C_6H_6 removed in a stream of air gives 3.7 g *p-nitrobenzenesulfonylhydrazide*, m $150-2^\circ$ (decompn), it is decompd by boiling EtOH . *Acetone p-nitrobenzenesulfonylhydrazide*, m $169-71^\circ$ (decompn), *benzaldehyde deriv*, m $142-4^\circ$ (decompn), *piperonal deriv*, m $189-90^\circ$ (decompn). The *m* isomer was prepd as above in 90% yield and m 130° (decompn) it is stable in boiling EtOH , *acetone m-nitrobenzenesulfonylhydrazide*, m $148-50^\circ$ (decompn). *MeEtCO deriv*, m $121-5^\circ$ (decompn). *benzaldehyde deriv*, m $150-1^\circ$ (decompn). *piperonal deriv*, bright yellow, m $173-5^\circ$ (decompn). *2,4-Dinitrobenzenesulfonylhydrazide*, prepd at -10° , light yellow, m 110° (decompn), it is very unstable to heat evolving N_2 heating with aq. NaOH or dil. HCl gives *2,4-dinitrobenzenesulfonic acid*, m 196° , either compd, heated in H_2O , gives quant. $\text{m-C}_6\text{H}_4(\text{NO}_2)_2$. The $\text{RSO}_2\text{NHNH}_2$ may be converted into RSOCl by passing Cl through a CHCl_3 soln. o and $\text{p-O}_2\text{N}(\text{C}_6\text{H}_4)_2\text{SO}_2\text{Cl}$ may be sepd by adding the soln in C_6H_6 to $\text{N}_2\text{H}_4 \cdot \text{H}_2\text{O}$ in abs. EtOH , heating at 35° , adding 0.5 N AcONa and crystg the $\text{p-O}_2\text{N}(\text{C}_6\text{H}_4)_2\text{SO}_2\text{NHNH}_2$, the AcONa soln contains the $\text{o-O}_2\text{N}(\text{C}_6\text{H}_4)_2\text{SO}_2\text{H}$, converted by Cl in CHCl_3 soln. o and $\text{p-O}_2\text{N}(\text{C}_6\text{H}_4)_2\text{SO}_2\text{Cl}$. Details are given for the sepn of the o and m , m and p , and o , m and p isomers, also for the sepn of 2,4,5- $\text{ClMe}(\text{O}_2\text{N})\text{C}_6\text{H}_3\text{SO}_2\text{H}$ from a mixt. of 2,4,5- and 2,4,6- $\text{ClMe}(\text{O}_2\text{N})\text{C}_6\text{H}_3\text{SO}_2\text{Cl}$. C. J. WEST

Tertiary arylalkylphosphines. **II.** IVOR K. JACKSON AND WM. J. JONES, *J. Chem. Soc.* 1931, 575-8, cf. *C. A.* 25, 283—2.5 $\text{Me}_2\text{C}_6\text{H}_4\text{PCl}_2$, b_p 129° , b_d 140° (20 g), and MeMgI gave 14 g of *p-xylyldimethylphosphine*, b_p 106° , d_4^{20} 0.9511; boiling in a current of air gives the oxide, m $91-5^\circ$, distils without decompn. *mercurichloride*, m 225° . CS_2 gives the inner anhydride of *p-xylyldimethylhydroxyphosphinecarbothionic acid* $\text{C}_6\text{H}_4\text{PMe}_2\text{CS}_2$ deep red, m 76° . MeI gives *p-xylyltrimethylphosphonium*

sulfide, m 204° , the *mercuriosulfide* m 152° , the *chloroplatinate*, orange, m 236° ; *trisulfide*, dark red, m 91° . *p-xylyldiethylphosphine*, b_p 157° , d_4^{20} 0.9302, *mercurichloride*, m 184° , the *methylphosphonium sulfide*, m 137° (*chloroplatinate*, orange, m 217° , *mercuriosulfide*, pale yellow m 105° , *trisulfide* chocolate, m 85°). *p-xylyldiisopropylphosphine*, b_p 161° , d_4^{20} 0.9281, *mercurichloride*, m 188° , *methylphosphonium sulfide*, m 105° (*chloroplatinate*, orange m 141° , *mercuriosulfide*, yellow, m 90° , *trisulfide*, m 76°). *p-xylyldibutylphosphine*, b_p 171° , d_4^{20} 0.9124, *mercurichloride*, m 179° , *methylphosphonium sulfide*, m 93° (*chloroplatinate*, orange, m 215° , *trisulfide*, m 70°). *p-xylyldisobutylphosphine*, b_p 184° , *mercurichloride*, m $163-6^\circ$, 172° , m 120° . *p-xylyldiamylphosphine*, b_p 214° , *mercurichloride*, m 117° , the *methylphosphonium chloroplatinate*, orange, m 151° . C. J. WEST

Magnesium derivatives of dichlorotriphenylphosphine and of the pentaphosphines. V. GREIGNARD AND J. SAYARD. *Compt. rend.* 192, 592-5 (1931)—Dichlorotriphenylphosphine (I), m 176° , was obtained by treating Ph_3PO with PCl_5 . I with RMgX gave $\text{Ph}_2\text{P}(\text{MgX})_2$ (II). II with ROH (R alkyl) gave Ph_2PR , m $163-6^\circ$, 172° , $179-82^\circ$ for R = Me, Et, Pr, resp. This is the 1st time that all 5 valences of P have been substituted by hydrocarbon radicals. JULIUS WHITE

Aromatic compounds with halogen in the benzene nucleus. K. A. KOBCHESKOV AND A. N. NESMEYANOV. *Ber.* 64B, 628-36 (1931), cf. *C. A.* 25, 927.—The following new compds were prepd. *Dichlorostannanes* ($\text{p-ClC}_6\text{H}_4)_2\text{SnCl}_2$ (I) (75% yield) by heating ($\text{p-ClC}_6\text{H}_4)_2\text{Hg}$ (II) (see *C. A.* 23, 5172) and SnCl_4 in abs. EtOH , m 86.5° , hydrolyzed by H_2O , *p-Br deriv* (III) (85% yield by a similar method), m 103° , *p-I deriv* (IV) (75% yield from ($\text{p-IC}_6\text{H}_4)_2\text{Hg}$ (V) and SnCl_4 in abs. Me_2CO), m 147° . *Dibromostannanes* ($\text{p-ClC}_6\text{H}_4)_2\text{SnBr}_2$ (92% yield from II and SnBr_4 in EtOH), m 73° (cf. Krause and Weinberg, *C. A.* 24, 349), *p-Br deriv* (96% yield), m 83° , *p-I deriv* (VI) (75% yield from V and SnBr_4 in Me_2CO), m 102° . *Diiodostannanes* ($\text{p-ClC}_6\text{H}_4)_2\text{SnI}_2$ (almost quant. yield from I and a large excess of NaI in hot abs. EtOH), m $46-7^\circ$, not hydrolyzed by H_2O , *p-Br deriv*, m $79-80^\circ$, *p-I deriv* (80% yield from IV and NaI in hot MeOH), m 88.5° . *Stannones* ($\text{p-ClC}_6\text{H}_4)_2\text{SnO}$ (90% yield from I in EtOH treated with an excess of 5% aq. KOH) sol in mineral acids giving the corresponding

salts: *p*-Br deriv. (92% yield from III), *p*-I deriv from IV. *Thioastannones*: (*p*-ClC₆H₄)₂SnS, prep'd by dissolving I in EtOH and adding a slight excess of 5% alc KOH sat'd with H₂S, *m* 170°, *p*-Br deriv, prep'd by passing H₂S into a soln of III in abs EtOH, *m* 228-9°, *p*-I deriv, prep'd by adding 10% alc KOH sat'd with H₂S to a soln of IV in abs EtOH, *m* 218°. *Trichloroastannanes* *p*-ClC₆H₄SnCl₃ (VII), prep'd in almost quant yield by heating I and SnCl₄ in a bomb tube at 150°, *m* 39°, insol in H₂O but made sol by the addn of a few drops of conc'd HCl, hydrolyzed by H₂O. *p*-Br deriv (VIII), by heating III and SnCl₄ in a bomb tube at 150°, *m* 64.5-5°, has properties similar to those of the corresponding Cl deriv. *p*-I deriv (IX), by heating IV and SnCl₄ in a bomb tube at 165°, *m* 55-6°, very hygroscopic, decomps in air in a few min. *Tribromoastannane* *p*-IC₆H₄SnBr₃ from VI and SnBr₄, *m* 80-80.5°, hydrolyzed by H₂O. *Stannonic acids* *p*-ClC₆H₄Sn(OH)₃, prep'd by dissolving VII in petroleum ether and adding an excess of 5% aq KOH, easily sol in aq caustic alkali and in mineral acids with formation of the corresponding salts. *p*-Br deriv, prep'd from VIII; *p*-I deriv, prep'd from IX, sol in mineral acids and caustic alkali. *Thioanhydrides of dithioastannonic acids*: (*p*-ClC₆H₄)₂Sn₂S₂, prep'd by dissolving VII in H₂O with the addn of a few drops of conc'd. HCl and satg the soln with H₂S, sol in HCl (1:1) with evolution of H₂S. *p*-Br deriv, prep'd from VIII, resembles the Cl deriv. *p*-I deriv, prep'd from IX, resembles the Cl deriv. Two compds contg multivalent I were prep'd (*p*-ClC₆H₄)₂SnCl₂, prep'd by dissolving IV in CHCl₃ at -15° and adding CHCl₃ sat'd with Cl at -15°, bright yellow, *m* 82-2.5°, *p*-ClC₆H₄SnCl₂, prep'd by passing Cl at -15° into as conc'd a soln as possible of IX in CHCl₃, bright yellow, decomps gradually 50-70°. L K

Piria reaction. L. The over-all reaction. W H HUNTER AND MURRAY M. SPRUNG *J Am Chem Soc* 53, 1432-43 (1931).—The Piria reaction is the prep'n of aminosulfonic acids from aromatic NO₂ compds by heating the latter with m tal sulfites and then boiling with mineral acids (*Ann* 78, 31 (1851)). In an attempt to explain the mechanism, a standard procedure has been developed for the reaction which consists in adding a weighed quantity of the NO₂ compd to enough of a 5.2 *N* soln of Na₂SO₃ to constitute approx a 75% excess over that calcd on the basis of 3 mols of bisulfite per mol of NO₂ compd. Enough 5 *N* NaOH was then added to neutralize 25% of the bisulfite and to the whole was added H₂O in the ratio of 120 cc for each mol of NO₂ compd. The mixt was then refluxed until homogeneous. In the acid treatment the mixt. was evap'd to about 50% of the original vol and treated while hot with 250 cc conc'd HCl for each mol of original NO₂ compd and boiled for 0.5-2 hrs. The combination of the 2 reactions gives in every case an aminosulfonic acid and an amine (except in the case of *p*-O₂NC₆H₄OH), it appears that a ring Me group favors the formation of amines, while the presence of a CO₂H, a 2nd NO₂, or a condensed ring favor the formation of the acids. In most cases the yield of the 2 chief products is slightly in excess of 80%. The speed of the 1st phase of the reaction (reduction) increased with diln. Using the standard procedure PhNO₂ gives 26% of PhNH₂ and 27% of *p*-H₂NC₆H₄SO₃H, *α*-C₆H₅NH₂ gives 2% of *α*-C₆H₅NH₂, 25% of naphthionic acid and 66% of 2,4-(HO₂S)₂C₆H₃NH₂, *o*-MeC₆H₄NH₂ gives 40% of *o*-MeC₆H₄NH₂, 31% of 2-toluidine 5-sulfonic acid and 12% of Na *o*-tolylsulfamate. *p*-MeC₆H₄NH₂ gives 72% of *p*-MeC₆H₄NH₂ and 11% of 4-toluidine 3-sulfonic acid. Nitro *p*-xylene (using 3 times the usual amt of H₂O) gives 60% of *p*-xyldine and 16% of 3-amino-1,4-xylene 6-sulfonic acid, *m*-C₆H₄(NO₂)₂ with a 10% excess of Na₂SO₃ gives a trace of *m*-O₂NC₆H₄NH₂ and 64% *m*-nitroamine-*p*-sulfonic acid, using a 110% excess of Na₂SO₃ gives 1% of this acid and 10% of *p*-phenylenediamine-*p*-sulfonic acid, *p*-O₂NC₆H₄CO₂H gives 21% of *p*-H₂NC₆H₄CO₂H and 61% of 4,3-H₂N(HO₂C)C₆H₃CO₂H, *p*-O₂NC₆H₄OH gives 14% of the same acid. These yields are the av. of several expts. Details are given for the isolation and identification of the reaction products. II. *The role of the sulfamic acids.* Ibid 1443-7.—The Na salts of phenylsulfamate (crystg with 1 mol H₂O), *p*- and *o*-tolylsulfamates and *p*-acetamidophenylsulfamate (crystg with 2 mols H₂O) have been prep'd, with HCl concn from 0.5 to 5.35 *N*, these salts hydrolyze quantitatively to H₂SO₄ and the corresponding amines, giving absolutely no trace of aminosulfonic acids. Hence the aminosulfonic acids which are produced by the 'over all' Piria reaction are certainly not formed by rearrangement of sulfamic acids. C J WEST

Preparation and properties of some ethyl arylsulfonates and a comparative investigation of their velocities of hydrolysis with those of the corresponding ethyl arylcarboxylates. LASLO DEMFNY *Rec trav chim* 50, 60-71 (1931).—The velocities of hydrolysis of PhSO₂Et and its *o*, *m* and *p*-nitro derivs have been measured and compared with those of the Et esters of BrOH and the 3-O₂NC₆H₄CO₂H. According to Wegscheider (*C A* 13, 2213 *Z physik Chem* 41, 52 (1902)) the hydrolysis of sulfonic esters by water is not accelerated by H ions and Olivier and Berger have shown (*C A* 17,

1783) that the same holds for all esters derived from strong acids. The difference in behavior between esters of CO_2H acids and of SO_3H acids clearly appears from an investigation of Karlsson (*C A* 20, 690) who showed the velocity of hydrolysis of AcOEt to be largely dependent on the H^+ ion concn. and to have a min. at pH 5.35, while the velocity of hydrolysis of PhSO_3Et (I) was nearly independent of the pH . The following compds. were prepd. by the interaction of the sulfonyl chloride and NaOEt in 1:10H: *Et benzenesulfonate* (I), $b.p.$ 153–4°, *F13 nitrobenzenesulfonate* (II), $m.$ 42°, *4 nitro isomer* (III), $m.$ 92°, *2 nitroester* (IV), $m.$ 15°, $b.$ 163–4° in an abs. vacuum. Working at 21° and with 30% EtOH (by vol.) as a solvent, the following values were obtained for K calcd. from the equation for a monomol. reaction (the values in parentheses represent the mean alkyl. of the soln.): I 0.00036 (0.005 N), 0.00041 (0.004 N), IV 0.0073 (neutral soln.), 0.0077 (0.0017 N), II 0.0022 (0.0017 N), 0.0023 (0.0035 N), 0.0025 (0.006 N), III 0.0026 (0.0012 N). For K calcd. from the equation for a bimol. reaction were obtained the following figures under the same conditions: I 0.46 (in water 1.96), IV 1.23, II 11.4, III 19.0 the alkyl. varying between 0.003 and 0.005 N . It now follows that the introduction of a NO_2 group into the benzene ring exerts an accelerating influence on the hydrolysis of both the sulfonates and the benzoates. While, however, a NO_2 group in the m and p positions has a similar effect in both types of esters, the effect produced by an o nitro group is very different, a slight accelerating effect being observed in EtOEt and a powerfully accelerating effect with I. With *Et p-toluenesulfonate*, $K = 0.00019$ (mean alkyl. 0.0017 N) was found under the same conditions, thus showing the retarding influence of a p -Me group. The difference between SO_3H and CO_2H acids is discussed at length in the original paper, to which reference must be made for details.

C F VAN DUIN

Arylsulfonalkylamides. LASLO DEMINY *Rec. trav. chim.* 50, 51–6 (1931), *cf.* *C A* 24, 835.—This paper deals with the arylsulfonalkylamides derived from PhSO_3H , $4\text{BrC}_6\text{H}_4\text{SO}_3\text{H}$, $p\text{-MeC}_6\text{H}_4\text{SO}_3\text{H}$ and $2,4,6\text{-Me}_3\text{C}_6\text{H}_2\text{SO}_3\text{H}$, which were prepd. from the corresponding sulfonyl chlorides and 2 mols. of the amines. They were purified by pptn. from 12% KOH with HCl , followed by recrystn. from a suitable solvent. *Benzenesulfonbutylamide* was obtained as a viscous oil which did not crystallize even at -30° , but was converted into a vitreous mass at that temp. (*cf.* Sselonina, *Chem. Centr.* 1899, II 809). *Amylamide* was also obtained as a viscous colorless oil, passing into a vitreous mass at -30° , *heptylamide*, $m.$ 20° , $4\text{BrC}_6\text{H}_4\text{SO}_3\text{H}$ was prepd. by sulfonation of PhBr with the addn. of a trace of I as a catalyst (*cf.* Rdy and Dey, *C A* 15, 513), yield 38 g. Na salt from 62 g. PhBr , the chloride $m.$ 75° , $4\text{Bromobenzenesulfonamylamide}$, $m.$ 55° , *hexylamide* $m.$ 55° , *heptylamide*, $m.$ 65° , $4\text{Methylbenzenesulfonbutylamide}$, $m.$ 43° , *omylamide* did not crystallize on standing for a long time in the ice chest, *hexylamide* $m.$ 62° , *heptylamide* $m.$ 27° , $2,4,6\text{-Me}_3\text{C}_6\text{H}_2\text{SO}_3\text{Cl}$, $m.$ 56° , was prepd. by adding 24 g. mesitylene with continuous stirring to 42 g. ClSO_3H at 0° ; yield, 20 g. $2,4,6\text{-Trimethylbenzenesulfonpropylamide}$, $m.$ 54° , *butylamide*, $m.$ 41° , *amylamide*, $m.$ 42° , *hexylamide*, $m.$ 64° , *heptylamide*, $m.$ 45° . From the graphs constructed with the aid of these m ps. it was noted in agreement with other investigators, that the introduction of alkylamine residues into arylsulfonalkylamides brings about a marked lowering of the m p. In general, however, a rise in the m ps. was observed with the introduction of the 5th CH_2 group and onward which was to be expected from Timmermans' law of the convergence of m ps. (*C A* 16, 2060). Considering the graphs obtained with 4 bromo- and with 2,4,6 trimethylbenzenesulfonalkylamides a min. in the m p. curves of the arylsulfonalkylamides is observed in the AmNH_2 series. In several other cases deviations or anomalous positions with regard to homologs have been noted with a normal alkyl group contg. 5 C atoms. This is found in the mol. vols. of the 5th term of the homologous series (Le Bas, *The molecular volumes of liquid chemical compounds*, 24, 146, 148, 150 (1915)), the velocity of hydrolysis of valericamide (Crocker, *C A* 1, 2076), the velocity of reaction of Amf on PhONa (*cf.* Segaller, *C A* 7, 3316, 8, 1267, Glueck and Kempf, *C A* 7, 3753) and the optical rotatory power of homologous series in which the 5th term shows an abnormal behavior (Pickard and Kenyon, *J. Chem. Soc.* 75, 362 (1899)). Thus, besides the ordinary oscillation (*cf.* Nekrasov, *C A* 22, 56, Verlaque, *C A* 24, 4204) there appears to be a 2nd type, which shows an anomaly in the 5th term in the homologous series (*cf.* van der Kam, *C A* 21, 2883).

C F VAN DUIN

New derivatives of p -arsanilic acid. I. p -Arsonosuccinilic acid and related compounds. GILBERT T. MORGAN AND ERIC WALTON, *J. Chem. Soc.* 1931, 615–9.—Compds. of the type $\text{H}_2\text{O}_2\text{AsC}_6\text{H}_4\text{NHCO}(\text{CH}_2)_n$, CONR_2 , from which the apparently toxic NHCH_3 group is absent, have been examd. *Atoxyl* (12.8 g.) and succinic anhydride (12 g.), heated 1 hr. at $170\text{--}80^\circ$, give 10 g. p -arsonosuccinilic acid (I), silky prisms, if the resulting mass is boiled with H_2O slightly acidified with HCl there results

succinamide *p*-*p'*-diarsonic acid, ill defined solid, *ds* Na salt, amorphous, dissolving in H_2O to give a soln of approx μ 7.5. I and PhNH₂ heated for 2 min give succinamide *p*-arsonic acid, small leaflets. Na salt, needles, giving a soln of approx μ 8; the methylamide deriv. prisms. Na salt, crystals with 2 mols H_2O , the ethylamide deriv., silky prisms yields a Na salt leaflets. I and HCl solid with SO₂ give *p*-dichloroarsino succinamide acid, pale buff m 210°. Heating I at 210° for 1 1/2 hrs. gives succinamid *p*-arsonic acid, amorphous, which regenerates I on addition of H_2O with 1 mol NH₃ at 75-80° for 1 1/2 hrs then results succinamidamide *p*-arsonic acid leaf to whose NH₂ salt crystallizes with 1 mol H_2O . Na salt, plates with 1 mol H_2O , the corresponding dimethyl amide forms leaflets. Na salt, crystallizes with 1 mol HCl the *p*-phenylene deriv. crystallizes as needles. Preliminary pharmacol reports on the mono-Na salts indicate very low toxicities combined with therapeutic activity against trypanosomes in all cases except that of the diarsonic acid. Values are given and compared with trypanazole.

C J WEST

Diphenyl ether series. III. Derivatives of the local anesthetic type. C M SUTER AND ELMER ORING. *J Am Chem Soc* 53, 1931, 9 (1931) of C 1 25, 1816—4 (4-nitrophenoxy)benzaldehyde (I) m 101.5 results in 75% yield from *p*-ClC₆H₄CHO and *p*-ClC₆H₄NO₂ after heating 6 hrs at 180°. the 2-Na salt m 84-5°. 4-O₂NC₆H₄OPh and AcCl with AlCl₃ in CS₂ give 52% of 4 (4-nitrophenoxy)acetophenone (II) m 80.1°. Oxidation of I or II gives *p*-Cl₂NC₆H₄OC₆H₄Cl, II m 215.1° chloride, m 79.80° amide m 167.8°, I ester m 71.5° reduction gives I 1 4-aminophenyl benzoate oil, the HCl salt decomps about 165° *lc* deriv m 122.1°. Bu 4-(4-nitrophenoxy)benzoate, yellow oil b 250.5° the NH₂ deriv yields a HCl salt, decomps 155-160°, *lc* deriv m 98.9°. 8 Diethylaminomethyl 4 (4-nitrophenoxy)benzoate HCl, decomps 100-110° (80% yield), NH₂ deriv, amorphous. Although these derivatives show considerable anesthetic action, they are too toxic to be of practical value. C J W

Substitution products of 2-nitro- and 2-acetamidodiphenyl ethers and the corresponding diphenylene oxides. HAMILTON MCCOMBIE, Wm Geo MACMILLAN AND HAROLD A SCARFROUGH. *J Chem Soc* 1931, 529-37—2 AcNH₂C₆H₄OPh and HNO₃ (d 1.42) at room temp for 15 min give the *Ac* deriv., yellow, m 140°, of 5-nitro 2-aminodiphenyl ether, yellow, m 116°, the diazo soln, dropped into hot 50% H₂SO₄ gives 2-nitrodiphenylene oxide, m 186°. More concn HNO₃ gives 5,4'-(2'-nitro 2-acetamidodiphenyl ether, pale yellow, m 190°, the free base, yellow, m 192°, *o*-ClC₆H₄OK and *o*-ClC₆H₄NO₂ heated 18 hrs at 170° give 2-ClC₆H₄OC₆H₄NO₂ 2, yellow, m 49°, the 2'-NH₂ deriv., pale yellow, b 203°, m 45° *lc* deriv., m 101°, an attempt to prep the corresponding diphenylene oxide was unsuccessful. HNO₃ gives the 5'-NO₂ deriv., yellow, m 142°, the free base, yellow, m 125° further nitration gives the *Ac* deriv., orange, m 176°, of 2-chloro-1-(2'), 5'-dinitro 2'-aminodiphenyl ether, yellow, m 202°, 2,5-Cl₂C₆H₃NO₂ and PhONa heated 5 hrs at 170-80°, give 4-chloro 2-nitrodiphenyl ether, pale yellow, b 211°, SuCl₂ in 1:60 HCl gives the 2-NH₂ deriv., b 215°, m 41°; HCl salt, m 192°, the *Ac* deriv. could not be purified, 3-chlorodiphenylene oxide, b 106°. Chlorination of 2-O₂NC₆H₄OPh or condensing *p*-ClC₆H₄ONa and *o*-ClC₆H₄NO₂ give 4-ClC₆H₄OC₆H₄NO₂ 2, m 46°, the 2-NH₂ deriv., b 215° HCl salt, m 181°, 4-chloro-2'-acetamidodiphenyl ether (I), m 99°, refluxing with Ac₂O and AcOH gives the *di* *Ac* deriv., m 106°. Nitration of I gives the *Ac* deriv., yellow, m 203°, of 4-chloro-5'-nitro-2'-aminodiphenyl ether, yellow, m 123°, there is also formed a small quantity of the 2-(2'), 5'-di-NO₂ deriv. of the *Ac* deriv., m 195°. 3-Chloro 5-nitrodiphenylene oxide, pale yellow, m 236°. *p*-ClC₆H₄OK and 2,5-Cl₂C₆H₃NO₂ give 4,2-Cl(O₂N)₂C₆H₃OC₆H₄Cl-4, m 75°, 2-NH₂ deriv., b 235°, m 66° HCl salt, m 145°, *Ac* deriv., m 106°, 3,6-dichlorodiphenylene oxide, b 190°. 4,2-Cl(AcNH)₂C₆H₃OC₆H₄Cl-4 on nitration gives the 5-NO₂ deriv., m 150°, and a small quantity of the 2-(2'), 5-di-NO₂ deriv., m 212°. Nitration of 4-BrC₆H₄OC₆H₄NHAc-2 or bromination of 5,2-O₂N(AcNH)₂C₆H₃OPh gives the *Ac* deriv., yellow, m 208°, of 4'-bromo-5-nitro 2'-aminodiphenyl ether, yellow, m 133°, 5,2-Br(11-N)₂C₆H₃OPh, through the diazo reaction, gives 2-bromodiphenylene oxide, m 120°. 5,2-Br(11-N)₂C₆H₃OC₆H₄Br-4 gives 2,6-dibromodiphenylene oxide, m 176°. 4,2-Br(11-N)₂C₆H₃OC₆H₄Br-4 gives the 7,6 isomer, m 195°, also obtained by bromination of (ClH₂)₂O in CS₂. 4,2-Cl(O₂N)₂C₆H₃OC₆H₄Br-4 results from the bromination of 4,2-Cl(O₂N)₂C₆H₃OPh or by the condensation of *p*-BrC₆H₄OK and 2,5-Cl₂C₆H₃NO₂. 2-O₂NC₆H₄OPh and ICl in AcOH give 4-iodo 2'-nitrodiphenyl ether, pale yellow, m 86°; Cl in CCl₄ gives the dichloride, m 96°. 4-IC₆H₄OC₆H₄NH₂ 2, b 210°, gives a HCl salt, m 219°, an *Ac* deriv., m 115°, and a *di* *Ac* deriv., m 91°. The results show that chlorination and iodination of 2-O₂NC₆H₄OPh follow the same path as bromination but that in these cases the introduction of more than 1 substituent atom is not attained.

C J WEST

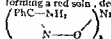
part of the I to sol products by CrO_3 , the amt of such oxidation depending upon the relative amt of oxidant used and the state of dispersion of I. Nevertheless when the above cycle of operations is repeated a large no of times a const. increase in insol material results. Therefore Green's conception of a definite oxidative condensation of I with 3 mols II must be abandoned.

Synthesis of simple and of substituted 2-alkylcinnamic alcohols, including a mono-molecular cubabin. MARSTON T. ROBERT AND GARFIELD POWELL. *J Am Chem Soc* 53, 1607-9 (1931).— $\text{PhCH}_2\text{CH}_2\text{CHO}$ (30 cc) is reduced by a soln of Mg in EtOH-HCl in a Natm giving 22 cc of 2 ethylcinnamyl alc, b_p 131-2° (cor), it instantly decolorizes Br in CCl_4 and is readily catalytically reduced to $\text{PhCH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$ and $\text{PhCH}_2\text{CH}_2\text{CH}_2\text{Me}$, *p*-nitrobenzoate, m 110-1° (cor) phenylurethan, m 63-4° (cor) 2 Amyl cinnamyl alc was similarly prep'd in 50% yield, it b_p 141-3°, phenylurethan, m 61-2° 2 Methyl 3 *p*-tolylallyl alc b_p 142-3° (60% yield) 2-Piperonylidenethanol (monomol cubabin) b_p 167-8° m 78-88° (cor), results in 20% yield with the Mg reagent and in 30% yield with Al(OEt)_3 benzoate, m 96-7° (cor), phenylurethan, m 96-7° (cor) C. J. WERT

The oxidation of benzyl alcohol by the three isomeric nitrochlorobenzenes. ROBERT E. LYONS AND GUIDO H. STEMPER, JR. *Proc Indiana Acad Sci* 38, 197-200 (1929).—L. *p*- $\text{C}_6\text{H}_4\text{ClNO}_2$ and PhCH_2OH . The reaction of PhCH_2OH with *p*- $\text{C}_6\text{H}_4\text{ClNO}_2$ gave BzOH and *p*-*p*-dichloroazobenzene, m 156°, with a little $\text{C}_6\text{H}_5\text{N}$. The reaction was initiated by dropping the soln of the nitro comp'd in xylene into a 78% NaOH soln contg the PhCH_2OH . The reduction at 139° was 100%, at 60°, 60.6%. The amt of NaOH used was not important. When the nitro comp'd was introduced in the solid form instead of as a xylene soln the product was 100% azoxy comp'd II. m $\text{C}_6\text{H}_4\text{ClNO}_2$ and PhCH_2OH . The reaction took place quantitatively at room temp yielding (m $\text{C}_6\text{H}_4\text{Cl}_2\text{N}_2\text{O}$ and PhCH_2OH III. *m*- $\text{C}_6\text{H}_4\text{ClNO}_2$ and PhCH_2OH . The conditions used successfully in the reactions of the *p*- and *m*- $\text{C}_6\text{H}_4\text{ClNO}_2$ with PhCH_2OH caused thermal decompn when the *o*-comp'd was used. By using a large amt. of solvent and dropping the alkali in slowly this decompn was curbed and a yield of 17% obtained. The product could not be identified but contained either a nitroso-, an azoxy or an azo group. H. M. STARK

The formation of thioamides from acylated aldehyde cyanohydrins. VII. JOHN F. OLIV AND TREAT B. JOHNSON. *Rec trav chim* 50, 72-6 (1931), cf *C A* 21, 98.—The practical method of synthesizing thiazoles is to allow an α halogenated ketone to interact with a thioamide. For the synthesis of certain thiazoles thioamides of HO acids were necessary and the latter were synthesized from acylated aldehyde cyanohydrins and H_2S in the presence of KSH , NH_4 , triethanolamine or pyridine. The superior action of these bases over NH_3 is ascribed by O and J to their greater basicity by virtue of which they are able to hold much larger quantities of H_2S in contact with the nitrile. The action of pyridine is not so favorable as that of the other bases mentioned above, which is not surprising in view of the fact that pyridine exerts a desulfurizing action on several org. compds (Raffo and Rossi, *C A* 8, 2683, 9, 1473). On account of its great cheapness and its remarkable influence on this addn reaction, triethanolamine was found to be the most suitable catalyst. The following compds were all prep'd on passing a current of H_2S through the EtOH soln of the acylated aldehyde cyanohydrin to which triethanolamine was added. Mandelic thioamide benzoate (cf Francis and Davis, *C A* 3, 2973), m 179°, acetate m 104°. Hydroxyacetic thioamide benzoate, m 103°, lactic thioamide benzoate, m 104°, α -hydroxybutyrylthioamide benzoate m 106°. The acylated aliphatic cyanohydrins could not be prep'd according to Francis and Davis' technique (*loc cit*) and were obtained by shaking the aldehyde, equiv amts of BzCl , NaCN and cracked ice until the odor of BzCl disappeared. The cyanohydrin was est'd with CHCl_3 or CCl_4 and purified by distn *in vacuo*. C. F. VAN DUIN

The nickel salt of benzamidoxime. J. V. DUNSET AND M. KURAS. *Chem Listy* 24, 454 (1930).—Benzamidoxime (1.3 g) in H_2O is mixed with 1.2 g $\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$ in H_2O and treated with 3 drops AcOH to prevent the pptn of Ni(OH)_2 . Dil NH_4OH is added until the odor of NH_3 is noticeable, and the mixt is left standing for 12 hrs. Fine, purple red needles are filtered off. The addn of NH_4OH to the mother liquor yields a 2nd crop of crystals. They are sol in hot H_2O , insol in cold H_2O , sol in EtOH , forming a red soln, decomp 90-125° and are stable in air. Analysis shows the structure



(I) In the presence of H_2O_2 the formation of I is hastened but no further oxidation occurs. With NH_4OH and NaOH an intense blue soln forms, boiling with PhO_2 decolorizes solns of I.

FRANK MARESH

Action of bromine on acetamidohydroxybenzoic acids and acetamidophenols. GUSTAV HELLER WITH IRITZ SÖLDNER *J prakt Chem* 129, 237-67 (1931).—5,2-AcNH(11O)C₆H₄CO₂H (I), m 218°, with 1 or 3 mols Br without or with heating on the water bath, gives the *mono-Br* deriv, m 217°. FeCl₃ gives a blue color, only with the excess of Br is a small quantity of bromoanil formed. I and HNO₃ in AcOH give a NO₂ deriv, darkens 270°, does not m 300°, 1 cCl₄ gives a violet color, heating with Br-AcOH gradually causes soln, giving bromoanil. 3 Acetamido-4 hydroxybenzoic acid (II), by reduction of the NO₂ deriv and acetylation, m 251.2° (decompn), 2 mols Br in AcOH, heated several hrs on the water bath, gives the *mono-Br* deriv, m 251° (decompn), FeCl₃ gives a brown color, 5 mols Br gives a *di-Br* deriv, m 219°. 1 cCl₄ gives no color, at a somewhat higher temp, there results a *tri-Br* deriv, pale yellow, m 230° (decompn), further heating of the reaction mixt gives *tribromo-2-bromocaminophenol* (7, III), yellow brown, darkens 180°, m above 305°. Concd HNO₃ and II give a 2,5,6-trinitro deriv, yellow, m 204°. 1 cCl₄ gives a dark red 1 cCl₄ reaction, with Br-AcOH this gives pentabromoacetanilide, m 222°, further heating gives pentabromoaniline, m 226°. 4 Acetamido-3 hydroxybenzoic acid, m 250.1°, 2 mols Br in 10 parts AcOH gives a *mono-Br* deriv, m 235° (decompn) 1 cCl₄ gives a blue color. 5 mols Br gives a compd apparently identical with III. 3 Acetamido-2 hydroxybenzoic acid, m 230° (decompn), 1 cCl₄ gives a deep blue color. 2 mols Br gives a *mono-Br* deriv, m 255° and giving an orange blue color with FeCl₃. 5 mols Br gives a *tri-Br* deriv, m 259°. 6,3 AcNH(11O)C₆H₄CO₂H and 2 mols Br give a *mono-Br* deriv, m 267°, 5 mols Br gives a compd, C₁₁H₉ONBr₂, carbonizes about 200°, which resembled III. 3 Acetamido-2,4-dihydroxybenzoic acid, m 220° (decompn) 1 cCl₄ gives a dark blue color. 2 mols Br in 15 parts AcOH gives a *mono-Br* deriv, decomps 240° and giving a deep blue FeCl₃ reaction, excess Br gives a mixt of products, which could not be sep'd. 6 AcNH(11O)C₆H₄CO₂H and 5 mols Br in 5 parts AcOH, on heating until HBr evolution ceases, give a mixt of the 5-Br, 3,5-Br and *tetra-Br* derivs, m 228°, further addn of Br to the reaction mixt and heating 20 hrs give BrC₆H₄NHAc in small yields. p-H₂NC₆H₄OH and Br-AcOH give bromoanil. p-AcNH(11O)C₆H₄OH and 5 mols Br in AcOH give the 2,6-di-Br deriv, m 188°, further heating of the filtrate gives bromoanil and tetrabromo hydroquinone. With 10 mols Br, there results the *tetra-Br* deriv, m 210-7°, heating with concd H₂SO₄ 3 hrs on the H₂O bath gives *tetrabromo-4-aminophenol*, m 215° (decompn) o-AcNH(11O)C₆H₄OH and 5 mols Br give a *tri-Br* deriv, m 163° (decompn), if the reaction mixt is heated 0.5 hr there results *tetrabromoacetylaminobenzene*, m 185°; if the latter is heated with 5 mols Br in AcOH for 4 hrs, there results a *penta-Br* deriv, m 207-8°.

C. J. WILKINSON

N-Methyl-N-phenylalkylaminoalkyl benzoates and p-aminobenzoates. ARTHUR C. CORB AND S. M. McILVAIN *J Am Chem Soc* 53, 1587-94 (1931).—Previous work indicated that it was desirable to prep and submit for pharmacol study a series of compds. of the procaine type in which various phenylalkyl groups would be attached to the N atom. This paper reports a series of compds. of the type BrO(CH₂)_xNMe(CH₂)_yPh, in which x is 2 and 3 and y is varied from 1 to 4. All b ps and m ps given are cor. γ-Phenylpropyl p-toluenesulfonamide, m 65.1-57°. δ-phenylbutyl deriv, m 53.5-39°. N-methyl-N-γ-phenylpropyl deriv, b₂ 23.1-8°, m 41.5-24°. N-methyl-δ-phenylbutyl deriv, b₂ 24.1-5°, m 60.5-11°. N-phenylpropyl p-toluenesulfonamide, m 113.3-37°. Hydrolysis gives the following products: 1. CH₃CH₂CH₂NHMe, whose HCl salt m 164.1-49°. PhCH₂CH₂CH₂NHMe, b₂ 55.6-1°, d₄²⁵ 0.9205, n_D²⁵ 1.50877; HCl salt, m 145.6-61°. δ-phenylbutylmethylamine, b₂ 95.0-54°, d₄²⁵ 0.9126, n_D²⁵ 1.50150; HCl salt, m 126.2-68°. These amines were condensed with p-O₂NC₆H₄CO₂CH₂CH₂Cl and γ-chloropropyl p-nitrobenzoate, b₂ 168.5-95°, d₄²⁵ 1.3222, n_D²⁵ 1.54756, the benzoates were isolated as the HCl salts in 50-75% yields, unreacted sec amines are removed by treatment with H₂Cl. The following N-methyl-N-phenylalkylaminoalkyl benzoate-HCl, BrO(CH₂)_xNMe(CH₂)_yPh, were prepd where x and y are: 2 and 1, m 145.6-64°, 2 and 2, m 134.2-48°, 2 and 3, m 106.3-7.1°, 2 and 4, m 106.9-7.5°, 3 and 1, m 145.4-61°, 3 and 2, m 128.5-94°, 3 and 3, m 117.5-8.3°, 3 and 4, m 121.7-5.7°. N-methyl-N-phenylalkylaminoalkyl p-nitrobenzoate-HCl, p-O₂NC₆H₄CO₂(CH₂)_xNMe(CH₂)_yPh, where x and y are: 2 and 1, m 216.6-7.6°, 2 and 2, m 170.9-1.9°, 2 and 3, m 122.6-3.6°, 2 and 4, m 120.6-1.6°, 3 and 1, m 206.6-7.4°, 3 and 2, m 147.2-8.2°, 3 and 3, m 99.5-100.3°, 3 and 4, m 159.3-60.3°. Catalytic reduction gives the N-methyl-N-phenylalkylaminoalkyl p-aminobenzoate-HCl, p-H₂NC₆H₄CO₂(CH₂)_xNMe(CH₂)_yPh, where x, y and z are: 2 and 1, m 210.6-10.6°, 2 and 2, m 245.2-7.4°, 2 and 3, m 182.3-4.3°, 3 and 1, m 219.1-20.1°, 3 and 2, m 190.7-2.2°, 3 and 3, m 178.8-9.8°, 3 and 4, m 156.8°. N-methyl-N-γ-phenylpropylamino-

ethanol, b_p 132.6-3.0°, d_4^{25} 0.9883, n_D^{25} 1.51733 (75.3% yield) *N*-Methyl-*N*- γ -phenylpropylamino-propanol, b_p 147.3-7.9°, d_4^{25} 0.9785, n_D^{25} 1.51335. The *N*-methyl-*N*-phenyl- α -alkyl benzoates are much less toxic than procaine, both subcutaneously and intravenously. The corresponding *p*-aminobenzoates show a decidedly different and most unusual toxic effect. Their intravenous toxicities are rather low as a group, certain ones approaching the toxicity of procaine. The subcutaneous toxicities, however, are considerably higher than would be expected from the intravenous values. *N*-Methyl-*N*-benzylaminethyl *p*-aminobenzoate is about 10% as toxic as procaine intravenously but is over twice as toxic subcutaneously. Details of the pharmacol. tests are given.

C. J. WEST

Titrimetric and spectrometric analysis of keto-enol mixtures. β -Phenylacetoacetic ester. KARL A. AUWERS. *J. Am. Chem. Soc.* 53, 1496-500 (1931). —Potemkin with Post and Michael, *C. A.* 25, 96. The rule that titrimetric and spectrometric details of the enol content of tautomeric mixts. gives results in satisfactory agreement holds for α -phenylacetoacetic ester (I). Titrimetrically, the enol content of the ester is found to be 39.2% and spectrometrically 36.7% therefore materially higher than for α -allylacetoacetic esters (II). For I, in contrast to II, the direct Br titration is just as applicable as the indirect Br titration.

C. J. WEST

Condensation of aromatic aldehydes with phenylacetimidate. THEODORA DE KIEWITZ AND HENRY STEIN. *J. Chem. Soc.* 1931, 639-40—24. (MeO) $_2$ C $_6$ H $_4$ CHO and PhCH $_2$ CN in EtOH LiONa give almost quant. 2,4-dimethoxy- α -phenylacetanitrile, pale yellow, m. 95°, it could not be demethylated. 3,4-di-MeO deriv., pale cream, m. 88°. α -Vanillin with 6 A. EtOH KOH gives almost quant. 8-methoxy-3-phenylcoumarin, straw-colored, m. 155.5°. 4,2-HO(MeO) $_2$ C $_6$ H $_4$ CHO with LiONa gives 4-hydroxy-2-methoxy- α -phenylacetanitrile, pale yellow, m. 105°, 4-Ac deriv., m. 158°, 4-hydroxy-3-methoxy deriv. pale cream, m. 99°, the yields of these 2 nitriles is 90% if 6 N EtOH KOH is used. All attempts to hydrolyze the nitriles were unsuccessful, they are thus probably *cis* forms.

C. J. WEST

Esculetin dimethyl ether from *Artemisia capillaris*. II. SHOICHI HARA AND CHUZO SHIMIZU. *J. Agr. Chem. Soc. Japan* 6, 1003-12 (1930), cf. *C. A.* 24, 5742. —Esculetin di-Me ether (I) was treated with concd. caustic alkali for several hrs. Dimethoxy-*o*-coumaric acid (II), m. 197.8°, was formed. Ac deriv. (III), m. 211°, was obtained from II. III was saponified by caustic alkali and then acidified with HCl. II was regenerated. I was demethylated by HIO $_4$ and esculetin, m. 268-70°, was obtained. I, m. 145°, was regenerated by methylation. Mono- (IV) and dibromide (V) (m. 235°) were prepd. from I in AcOH or benzene with Br. Sapon. IV by Tilden and Burrow's method colorless needles, m. 245-6°, were obtained. They were dimethoxycoumaric acid. Colorless needles, m. 220-2°, (MeO) $_2$ C $_6$ H $_4$ BrCO $_2$ H were obtained from V. One of the Br atoms may be combined with the side chain and the other with the benzene nucleus. Hence the formula of V should be C $_9$ H $_5$ O $_4$ Br $_2$, not C $_9$ H $_4$ O $_4$ Br $_2$. Y. KIMURA.

Action of hydroferrocyanic acid upon bicyclic terpenes; a new partial synthesis of terpene bases and alcohols. K. STEPHAN AND TH. HAMMERICH. *J. prakt. Chem.* 129, 283-305 (1931). —Camphene and H $_2$ Fe(CN) $_6$ give the compd. (C $_9$ H $_9$) $_2$ Fe(CN) $_6$. Heating 160 g. of this compd. in 12 l. 15% KOH in an autoclave at 160° (9-10 atm.) for 13 hrs. gives 7% isobornylamine, 60% nearly pure camphene and 43% of a mixt. contg. 74.4% camphene, 23.1% camphene hydrate and 2.5% isoborneol. Details of the sepn. of these products and their identification are given. The compd. with α -pinene on alkali decompn. at 160° gives 70% of a mixt. of α -terpineol, limonene and dipentene and 30% of a mixt. of isobornyl and isofenchylamines. The verbenol formed from α -pinene in the air also forms a complex with H $_2$ Fe(CN) $_6$, which is easily decomposed with the reformation of the verbenol. A terpineol formed by the action of H $_2$ O on α -pinene behaves in the same way. Nopinene (8-pinene) also gives a mixt. of compds. with H $_2$ Fe(CN) $_6$, the easily decomposed part gives α -terpineol and terpene hydrate, the more difficultly decomposed part gives 40% of isofenchylamine and 55% of α -terpineol and limonene.

C. J. WEST

Fixation of hydrogen by acetylene derivatives. XVIII. YU. S. ZALKIND AND V. O. MOKHACH. *J. Russ. Phys. Chem. Soc.* 62, 1643-7 (1930). —Dihydroxybornyl-acetylene C $_9$ H $_9$ (OH) $_2$ CC $_2$ H $_2$ OH, was synthesized by mixing 90 g. camphor in 200 cc. ether with 1 C $_2$ H $_2$ Br $_2$, prepd. from 18 g. Mg, 90 g. 1 Br and 230 cc. ether. The mixt. was left for several days, being heated 4 or 6 hrs. daily or altogether about 20 hrs. After decompn. with water a basic Mg salt was sepd. and dissolved in 20% AcOH, the product, extd. with Et $_2$ O and recrystd. from petr. ether, m. 231-2°, is sol. in Et $_2$ O, acetone, CHCl $_3$, benzene. Yield 43%, analysis confirmed the compn. The assigned

ure was proved by heating with KOH which decomd the compd into camphor $1\text{C}_{11}\text{H}_{17}$. For H fixation several catalysts were tried. Pd black is more efficient than loidal Pd and that pptd on BaSO_4 . Pd black (0.6 g) effects an addn of 2 H atoms the mol of the glycol (1.6 g dissolved in 50 cc Et_2O) in 10–15 min while the satn. required 315 min. This difference in time of reaction is characteristic for hydrogenation a triple bond with Pd, as Pt requires for fixation of the first 2 atoms as much time as the 2nd pair. The ethylene glycol crystd from $\text{MeOH} + \text{Me}_2\text{CO}$ gives crystals 165–7° sol in ether, benzene, CHCl_3 , petr ether, mol wt 316–323 (calcd 332), es a yellow color with coned H_2SO_4 , does not decolorize KMnO_4 , but reacts with Br_2 a satd glycol, recrystd from acetone, m 201°, sol in CHCl_3 , ether, benzene, petr ether, mol wt 315 (calcd 311) the yellow color with H_2SO_4 is very slight, Br and KMnO_4 are not decolorized. This acetylene glycol is more difficult to synthesize and to drogenate than other acetylene glycols.

J. G. TOLPIN

Action of substituted aromatic amines on camphoric anhydride. The rotatory powers of some disubstituted camphoranic acids. MAHAN SINGH AND DARA SINGH *Chem Soc* 1931, 478–83, cf *C A* 24, 4286.—The following substituted camphoranic acids were prepd by heating camphoric anhydride and the amine with fused AcONa at 5–50° for 3–4 hrs, in no case was there any residue of imide. 2',3'-di Me (I), m 1–2°, 2',6'-di Me (II), m 236–8°, 2',5'-di Me (III), m 203–4°, 5'-nitro 2'-methoxy (IV), m 162–3°, 4'-nitro 2'-methoxy (V), light brown, m 185–6°, 5'-nitro 2'-methyl (VI), straw colored m 220–1°, 3'-nitro 1'-methyl (VII), light brown, m 204–5°, nitro-4'-methyl (VIII), deep yellow m 187°, 4'-nitro 2'-methyl (IX), m 220–30°, the following acids were prepd by the action of fuming HNO_3 in AcOH . 4'-chloro-2'-nitro (X), deep yellow, m 204–5°, 4'-bromo 3'-nitro (XI), deep yellow, m 212°, 2',6'-nitro-4'-methoxy (XII), light orange, m 228–9°, 2',6'-dinistro 4'-ethoxy (XIII), orange, 180–90°. Mol rotatory powers in MeOH , EtOH , Me_2CO and MeCOEt , resp., retd as follows. 2'-Me 151, 141, 98, 91, 1168, 149, 120, 109, III 203, 165, 147, 127, 174, 168, 109, 102, 2'-MeO 30, 28, –16, –10, 4'-MeO 163, 153, 123, 100, 4'-EtO 0, 154, 110, 104, IV 141, 128, 109, 101, V feeble rotation XII –306, –412 –270, 255, XIII –263, –373, –211, –192, 4'-Me 170, 146, 122, —, VI 214, 219, 154, I, IX 96, 80, 77, —, VII 141, 123, III 101, VIII –167, –202, –161, –138. Two cups of the same polarity reinforce each other when they are in the β position with respect to each other, e.g., III and VI exceed any of their resp isomers in rotatory power. Pairs of opposite polarity neutralize each other's effect. The rotatory power of IV practically the same as that of the unsubstituted compd. The NO_2 group in the 4'-position has a depressing effect on the rotatory power of the original compd.

C. J. WEST

Stereochemistry of biphenyl compounds. XIV. Preparation and resolution of 3,3',5,5'-tetramethyl-2,2'-difluoro-6,6'-diaminobiphenyl. I. C. KLUDERER AND J. R. ADAMS, *J Am Chem Soc* 53, 1575–60 (1931), cf Boek and A. C. A. 25, 1–3, 5,2,6-Me₂(H₂N)(O₂N)C₁₁H₇Br was transformed into the 2 diazonium borofluoride (7% yield), decomps about 195°, which decomps on heating to 3,5-dimethyl-2-fluoro-nitro-1-bromobenzene, yellow, m 49–51° (45% yield), with Cu in PhNO_2 there results 3,5,3',5'-tetramethyl-2,2'-difluoro-6,6'-diaminobiphenyl, m 231–6°, reduction with Zn and AcOH gives 64% of the 6,6'-diamino deriv (I), m 154–5° (cor.) resolution of I was effected by d -camphorsulfonic acid in abs AcOH , d I, m 151–2° (cor.), $[\alpha]_D^{20} -4.1^\circ$ (0.4000 g in 15 cc Me_2CO) d -camphorsulfonate, m 171–3°, $[\alpha]_D^{20} 0^\circ$ (0.1500 g in 15 cc abs EtOH), d I, m 150–1° (cor.), $[\alpha]_D^{20} 3.2^\circ$ (0.1000 g in 15 Me_2CO); d -camphorsulfonate, m 125–30°, $[\alpha]_D^{20} 32.5^\circ$ (EtOH), the active forms are readily racemized by heating in a neutral solvent, such as MeOH or EtOH or more readily in glacial AcOH . 3,5-Dimethyl-1-bromobenzene-2-diazonium borofluoride, compd 161° (60% yield), on heating, gives quant 3,5-dimethyl-2-fluorobromobenzene, m 87–9°, $d_{40}^{20} 1.452$, $n_D^{20} 1.3100$, nitration gives 88% of the 4,6-di- NO_2 deriv, light flow, m 80–7°, Cu in PhNO_2 gives 69% of 3,5,3',5'-tetramethyl-2,2'-difluoro-4,6,4',6'-ranitrobiphenyl, m 202–4° (cor.), the *trans*- NH_2 deriv, yellow, m 250–3° (74% yield) is could not be resolved by d -camphorsulfonic acid, d -bromocamphorsulfonic acid or tartaric acid.

C. J. WEST

Monoacetylation and monodiazotization of diaminobiphenyl. C. F. GAZZ *Am Ital* 61, 31–42 (1931).—In connection with an investigation in progress on the epn of derivs of the Ph_2 series, a special study was made of the behavior of o - $\text{H}_2\text{NC}_6\text{H}_4\text{C}_6\text{H}_4\text{NH}_2$ (I) when monoacetylated and monodiazotized in order to establish which the 2 NH_2 groups first reacts with AcOH and with HNO_2 , resp. A survey of the literature shows that most Ph_2 derivs known up to now are sym, but since I is unsym.,

it is only by settling the initial reaction mentioned above that the constitution of I derives obtained by the Sandmeyer reaction can be established. I (5 g) in hot 60% EtOH (50 cc) and Ac₂O (250 g) brought almost to boiling, poured into cold water, let stand several days, decanted, the residue (a mixt. of a cryst. compd. and a pitch) washed with 1 t.O. (to remove the pitch), the cryst. residue digested repeatedly with a large vol. of warm dil. HCl (the insol. residue is the di. Ac deriv., cf. Ann 207, 331), filtered, the filtrate neutralized with Na₂CO₃ and the ppt. purified by repeated recrystns. from dil. EtOH and decolorization with animal charcoal, yields 2-amino-4'-acetylamino-phenyl (II), m 170-7°. I (10 g), dissolved in hot 40% H₂SO₄ (50 cc), cooled (with rapid agitation to assure the formation of fine crystals) to -4° or -5°, aq. NaNO₂ (37 g in 15 cc) added, let stand at room temp., heated until N is no longer evolved, filtered, the filtrate decolorized with animal charcoal, made alk. with aq. Na₂CO₃, the ppt. dissolved in 5% aq. KOH, filtered, acidified with excess of HCl, filtered, decolorized again with animal charcoal, reprecip. with Na₂CO₃ and recrystd. from boiling EtOH, yields 2-hydroxy-4'-aminobiphenyl (III), already described but the exact constitution not established by Bamberger (Ann 390, 161), yellowish, m 141-2°, is readily diazotized at room temp. with formation on heating of 2,4'-dihydroxybiphenyl (cf. Ber 13, 224, Ann 207, 357). II (5 g), suspended in 20% H₂SO₄ (10 cc), diazotized at 0° with aq. NaNO₂ (2 g), let stand a short time, heated gently until N is evolved, the ppt. washed with dil. HCl (to remove small quantities of III), and purified by repeated recrystns. from boiling EtOH, yields 2-hydroxy-4'-acetylamino-phenyl (IV), pale yellowish, m 198-0°, sol. in aq. alk. hydroxides and insol. in dil. acids. Ac₂O (2 cc) added to III (1 g) (spontaneous heating), let stand a short time, poured into water, let stand until all excess Ac₂O has dissolved, filtered, the residue washed with water and recrystd. from EtOH, yields IV, III (2 g) and Ac₂O (7-8 cc), boiled for a short time, poured into cold water, the pitchy ppt. washed with Et₂O, the residue washed with dil. NaOH, then with HCl, and crystd. from boiling EtOH, yields 2-acetoxy-4'-acetylamino-phenyl, m 138°, aq. NaNO₂ (0.63 g), added to II (2 g) suspended in ice-cold water and coned. HCl (28 cc), 50% H₃PO₄ (15 cc) added and the ppt. recrystd. from boiling EtOH, yields 4-acetylaminobiphenyl (V), m 170°. If the reduction is carried out with EtOH, a considerable yield of IV and of pitchy substances are also formed. V and excess dil. HCl, boiled a long time, dil. with hot water, cooled, diazotized with NaNO₂, heated almost to the b. p. until N is no longer evolved, and the ppt. recrystd. from EtOH, yields 4-hydroxybiphenyl (VII), m 160° (the highest m. p. in the literature is 164-5°). VII can also be prepd. thus: I (5 g), dissolved in hot 40% H₂SO₄ (40 cc), cooled to -4° and to -5°, aq. NaNO₂ (1.8 g in 10 cc) added, EtOH (equal vol.) added, let stand overnight, the soln. sep'd. from the pitch, the EtOH evapd. *in vacuo*, the residue made alk. with aq. Na₂CO₃, the ppt. (mixt. of 4-aminobiphenyl (VIII) and III) dissolved in 5% KOH, the insol. residue (VIII) dissolved in dil. HCl, decolorized with animal charcoal, diazotized with NaNO₂, the diazo deriv. decomposed by heating, and the ppt. recrystd. from EtOH, yields VII, m 160-1°. III (2 g) in dil. HCl, diazotized at room temp., 50% H₃PO₄ (15 cc) added, warmed gently until N is no longer evolved, the sep'd. oil washed with water and crystd. from petr. ether, yields 2-hydroxybiphenyl, m 56°. Other reducing agents were tried, but in all cases the yields were very low.

C. C. DAVIS

Oximes of o-hydroxybenzophenone. E. P. KOHLER AND W. F. BRUCE. *J. Am. Chem. Soc.* 53, 1569-74 (1931).—The stereoisomeric oximes of o-HOC₆H₄COPh cannot be dehydrated to a cyclic compd. Their behavior, therefore, contributes no independent evidence with respect to the proper interpretation of the Beckmann rearrangement, but indicates that this rearrangement is reliable when it is employed with sufficient care. Both of the oximes may be obtained directly from the ketone and NH₂OH in alk. soln., at the ordinary temp. the product is almost pure *k*-oxime (I), in boiling solns., the product is a mixt. in which the quantity of *n*-oxime (II) increases with the time that the soln. is boiled. I crystallizes from C₆H₆ in thin plates, m 142-3°, II is probably obtained most readily by heating I with strong alkalis, II crystallizes in needles, m 141-2°. II is stable to boiling 20% aq. NaOH for 2 hrs. I sublimes at 175-80°, while II is changed into I at 185-200°, but even at 185-200° the process is exceedingly slow. With PCl₅, I gives o-HOC₆H₄CONHPh, while II gives 2-phenylbenzoxazole (III). Attempts to dehydrate the oximes with dry HCl were unsuccessful. I gives a yellow HCl salt, m 118-20°, decomps. with EtOH gives I but heating at 140-5° gives III and the ketone, II also gives a HCl salt, m 132-4°, which likewise gives III on heating. HCO₂H changes II into I and also causes a Beckmann rearrangement.

C. J. WEST

The oxidation of acetylene glycols. An o-diketone of the tetrahydrofuran series. T. I. TEKNIKOVA AND P. A. TIKHOMOLOV. *J. Russ. Phys.-Chem. Soc.* 62, 1217-22

(1930) — *Tetralolylbutenediol* (5 g) in 25 cc AcOH with 2.2 g CrO_3 in 25 cc. 95% AcOH gave 2,2,5,5-tetralolyl-3,4-diketotetrahydrofuran (I), and 2 g of itolyl ketone (II) was obtained in 42% yield as crimson red crystals, m 182° , sol in PhH, CHCl_3 , and AcMe, almost insol in AcOH and EtOH. At $50-60^\circ$ I is oxidized to II by CrO_3 in AcOH. I gives a phenylhydrazone, orange yellow, m $181-6^\circ$, an oxime, pale yellow, m 220° (decompn.), and a deriv with $\text{C}_6\text{H}_5(\text{NH}_2)_2$, m $230-1^\circ$. Only one CO group reacts with NH_4OH or PhNHNH_2 . I with MgI and EtMgBr gave rose crystals, m 150° , extremely labile and oxidizing rapidly back to I. I was not attacked by PhMgBr nor Zn and HCl .

Lewis W. Butz

The action of halogen acids on acetylene glycols. VI. The action of hydrogen iodide on dimethyldiphenylbutenediol. YU S ZALKIND AND S V NEDZVETSKII. *J. Russ Phys Chem Soc* 62, 1011-19 (1930), cf *C A* 21, 56. — Symmetrical dimethyldiphenylbutenediol (isomer m 143°) (I) gives with III a mixt of 2,5-dimethyl-2,5-diphenyl-3-iodo-2,5-dihydrofuran (II) and 2,5-diphenyl-3,4-diiodo-2,4-hexadiene (III). I with a satd soln of III in AcOH (10 cc per 2 g of glycol) at room temp for 4 hrs gave a reaction product sepg as a heavy oil when poured into H_2O . The oil was washed by decantation, dissolved in H_2O , washed with $\text{Na}_2\text{S}_2\text{O}_3$, and dried with CaCl_2 . Vacuum distn after removal of the H_2O gave a yellow oil in 80% yield, easily sol in AcMe, Et_2O and PhH, more difficultly in EtOH and petr ether. The oil was not further investigated, but the analytical results and mol-wt detn indicated II mixed with III. I with III in AcOH (10%) at room temp after 2.5 hrs gave crystals of III in 25% yield, colorless, m $145-6^\circ$, very sol in AcMe, Et_2O , PhH, hot EtOH and petr ether. The filtrate from III contains II and unchanged I. Increased III concn increases the amt of II. III does not decolorize KMnO_4 in the cold, does not add Br_2 , and is not saponified by heating 20 hrs with 20% KOH. III with 3% KMnO_4 in AcMe at $40-5^\circ$ or with CrO_3 in AcOH at the b p is converted to a small extent to AcPh and a small amt. of acids not identified. The isomer of I which m. 126° (IV) gave with 10% III in AcOH at room temp after 2.5 hrs 20% of III together with some I. IV is isomerized to I by III. The reverse transformation by III was not observed. Probably only I gives II and III directly. III (8.5 g), refluxed 9 hrs with 100 g EtOH and 6 g Zn dust, gave V, m 84° , easily sol in AcMe, CHCl_3 , EtO, PhH, EtOH, and petr ether. Analysis of V corresponded roughly to $\text{C}_{11}\text{H}_{11}\text{I}$. V in the cold decolorizes neither Br_2 in CHCl_3 nor alk KMnO_4 soln. V did not split out III with alc KOH nor was it oxidized by heating 8 hrs at the b p with 3% KMnO_4 in AcMe. The analysis for V indicated a compd richer in H than required by the suggested formula, but since V is so stable to oxidants it is not an olefin. Heating III in H_2O with Mg or in PhH or cumene with Zn leaves it unaltered. Heating 3 g III, 1.5 g Zn, and 15 cc of freshly distd quinoline for 1.5 hrs. at the b p in H_2O or N_2 gave VI, m $103-4^\circ$, sol in AcMe, CHCl_3 , PhH, petr ether and EtOH. VI is probably $\text{C}_{11}\text{H}_{11}\text{O}$, the structure was not detd.

Lewis W. Butz

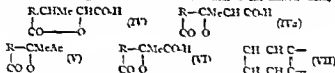
Hydrogenation of certain branched compounds over nickel. HOWARD ADKINS, WALTER H. ZARTMAN AND HOWARD CRAMER. *J Am Chem Soc* 53, 1423-8 (1931). — Exptl conditions for the successful hydrogenation over Ni of Ph_3COH , Ph_3CH , $\text{dicyclohexylphenylmethane}$ (I), 1,3,5- $\text{C}_6\text{H}_3\text{Ph}_3$, Ph_3N , 1,3,5- $\text{C}_6\text{H}_3\text{Me}_3$ and (2,4,6- $\text{C}_6\text{H}_3\text{Ph}_3$) are reported. H_2O and EtOH inhibit the hydrogenation of I. 1,3,5-Tricyclohexylcyclohexane, b_p , 228-8.5°, m $159-60^\circ$. 2,2',4,4',6,6'-Hexamethylbicyclohexyl, b_p 123-6°, d_{25}^{25} 0.8932, n_D^{25} 1.4873. Tricyclohexylmethane, b_p 164-5°, m $58.5-9.5^\circ$; tricyclohexylamine, b_p 188-9°, m $160-1^\circ$. HBe salt, m $267-8^\circ$, HCl salt, m 264° , picrate, m $172.5-3^\circ$.

C. J. West

Mixed benzoin. III. The structure of some unsymmetrically substituted desoxybenzoins. JOHANNES S. BUCK AND WALTER S. IDE. *J. Am Chem Soc* 53, 1536-12 (1931), cf *C A* 24, 5748. — The Beckmann transformation has been used to det the structures of certain unsymmetrically substituted desoxybenzoins and to assign configurations to the oximes derived from them. Desoxy compds. of the mixed benzoins formed from the following pairs of aldehydes were investigated: $\alpha\text{-ClC}_6\text{H}_4\text{CHO}$ and veratric aldehyde (I), $p\text{-MeOC}_6\text{H}_4\text{CHO}$ (II), $p\text{-MeNC}_6\text{H}_4\text{CHO}$ (III), piperonal (IV), BzH and $p\text{-MeOC}_6\text{H}_4\text{CHO}$ (V), piperonal (VI) and $p\text{-MeNC}_6\text{H}_4\text{CHO}$ (VII). Reduction of I gives $\text{ClC}_6\text{H}_4\text{CH}_2\text{COC}_6\text{H}_4\text{OMe}$ (I), whose *anti-oxime* m 137° (64% yield) and yields on the Beckmann rearrangement 1-chlorophenylacet-3,4-dimethoxyanilide, m 177° ; this was also synthesized by heating the acid and amine at $180-200^\circ$ for 2 hrs. II gives $\text{ClC}_6\text{H}_4\text{CH}_2\text{COC}_6\text{H}_4\text{OMe}$ on reduction, whose *anti-oxime* m 97° (86% yield), rearrangement gives 1-chlorophenylacetanilide, m 163° . Reduction of III gives 1-chlorobenzyl 4-dimethylaminophenyl ketone, m 122° ; *anti-oxime*, m 173° , rearrangement gives 1-chlorophenylacet-4-dimethylaminoanilide, m 165° , also synthesized by the

Schotten-Baumann reaction. IV gives 1,4-dibromo-2,4-methylenediphenyl ketone in 105% *ortho-oxime* in 121° (42% yield) rearrangement gives 1-chlorophenyl acryl-2,4-methylenediphenyl ketone, bp. m 175°. V gives 4-MeOC₆H₄CH₂COPh *ortho-oxime*, m 133 (23% yield) rearrangement gives 4-methylphenyl ketone in 117% *ortho-oxime*, m 94° (10% yield) rearrangement gives 4-MeOC₆H₄CH₂NHPh, m 9°. PhCH₂COCl and PhOMe give benzyl 4-methoxyphenyl ketone in 73% *ortho-oxime* in 114° (4% yield) rearrangement gives phenyl ketone in 121°. Reduction of VI gives (64% of benzyl 4-methoxyphenyl ketone in 4% *ortho-oxime*, m 113° (60% yield) rearrangement gives phenyl ketone 1,4-methylenediphenyl ketone, bp. m 141°. Reduction of VII gives PhCH₂COCH₂CH₂Me the *ortho-oxime* gives on rearrangement phenylacetyl-dimethoxybenzyl ketone cream, m 144°. The above results support the view that the transformation takes place between the vicinal groups of the oximes. C. J. West.

Congo copal oil. II. The oxidative degradation of the naphthalene hydrocarbon C₁₈H₁₄ from Congo copal oil. L. WESTENBERG AND J. P. WILBERT. *Exp. tar chem.* 50, 188-94 (1931) cf. C. A. 23, 3214.—The same hydrocarbon C₁₈H₁₄ was obtained by Westenberg (C. A. 23, 3216) on dehydrogenation of Congo copal oil with S. by Kuzicka, Steiner and Schurz (C. A. 23, 988) by pyrolytic decomposition of a mixt. of acids from Manila copal, followed by dehydrogenation with D and by Harver, Heilbron and Kamm (C. A. 21, 1112) on dehydrogenating squalene with S. The structure of this hydrocarbon as 1,2,5-C₁₈H₁₄Me was proved by Kuzicka and Hocking (C. A. 25, 1232) and Heilbron and Wilkinson (C. A. 25, 654). On oxidizing the hydrocarbon in the way described by Westenberg (*loc. cit.*) with CrO₃, the reaction products, sol. in NaHCO₃, do not only contain an acid C₁₈H₁₂O₄, but also the ketonic acid 2,6-MeAcC₆H₃CO₂H (I), which can be isolated by boiling out the NaHCO₃-sol. reaction products with water and reprecip. the only residue several times from ligroin and water. 1-methyl-2-acetyl-2-benzoyl, acid (I), m 116° is then obtained in 10% yield *oxime*, m 162°. I may be converted easily into its *Verzine* m 67-8, thus showing that the so-called esterification rule of Victor Meyer is only to be used with caution. The oxidation with KMnO₄ converts I into hemimellitic acid (m. p. of the anhydride 102° and mixed with an authentic specimen (m. p. 135°) prepd. from 1-C₁₈H₁₄(CO₂H)₂, m 193°). With NaOH, I is converted into 3-methylphenol, acid (II) (m. p. of the anhydride 111°) identical with an authentic specimen prepd. from m-toluic acid (Jurgens, C. A. 2, 470; Muller, C. A. 3, 1018). The oxidation of the hydrocarbon C₁₈H₁₄ with less CrO₃ gives *isomellitic β naphthoquinone* (III) (R = VII) in 1% yield with C₆H₅(NH)₂ this quinone forms a *quinoxaline* m 142° while the oxidation with CrO₃ gives the acid IV or Va, m 234°, already obtained by Westenberg (*loc. cit.*). The lactone character of the latter acid was proved by the fact that the exact neutral soln. became alk. in the course of a few min. The oxidation with CrO₃ or KMnO₄ converts IV into an acid C₁₈H₁₂O₄, m 235-6°, which was not investigated further. The *lactone Verzine* (V), m 61°, was prepd. according to Westenberg (*loc. cit.*) *oxime* m 116°. The oxidation of V with NaOH, with H₂O₂ in alk. soln. or with KMnO₄ gives the acid (VI), which contains 1 H₂O and m 74°, on heating at 100° the water of crystn. is given off with the formation of the anhyd. acid, m 130°.



C. F. VAN DUSEN

The dibromonaphthalenes obtained by the action of bromine upon naphthalene, YU. S. ZALKIND AND S. R. FALKMAN. *J. Russ. Phys. Chem. Soc.* 62, 1121-32 (1931).—Earlier workers did not consider the influence of reaction conditions upon the yields and nature of di Br isomers obtained by the direct bromination of C₁₀H₈ (I). One procedure used in the present work follows: With strong reech. agitation in powder I is added to excess Br₂ in 1.6 g. NaOH. This is followed by the addn. of 1 N HCl when crystals sep. These are washed successively with water, warm NaOH and water, and the product is then recrystd. from EtOH giving 2 fractions A, m 81-2° and B, m 67-8°. The milder conditions used e. g. taking the I and Br₂ in suitable solvents (CS₂ or CHCl₃), slow addn. cooling etc. the greater the yield of A, which is shown conclusively to be 1,4-dibromonaphthalene (II). Addn. of Fe does not affect the yield of II if mild conditions are employed, but at high temps. with undissolved I, the yield of II is greatly reduced in the presence of Fe. Py bromination of I at 60-70° with Br₂ in alk. soln., a theoretical yield of B is obtained. The highest yield of II obtainable directly from the

bromination mixt. was 106° . B was found to be a eutectic mixt. of II with 1,5-dibromonaphthalene (III), m. 131° , in the ratio 3:1. This eutectic mixt. is very difficult to resolve into II and III, although by further bromination II can be converted to $C_{10}H_6Br_2$, Br₂ and $C_{10}H_4Br_2$, while III is not affected. It was found possible to obtain III in 10% yield by treating α - $C_{10}H_7Br$ at $170-80^{\circ}$ for 13 hrs. with excess Br₂, and recrystallizing successively from petroleum ether and EtOH. II and III were the only dibromonaphthalenes obtained by direct bromination. Lewis W. Buttz

Derivatives of 1-chloro-2-nitronaphthalene. A. P. J. HooGVEEN. *Rec. trav. chim.* 50, 37-40 (1931). Previously II showed that 1,2- $C_{10}H_6ClNO$ is readily produced by the action of iodine on powder soln. of β - $C_{10}H_7NH_2$ (C. A. 25, 914). Backer and Klavars, C. A. 25, 915, and this procedure was used for the prepn. of a larger amt. of this compound, details of which are given in the original paper. On heating the EtOH soln. with concd. HCl this compound was hydrolyzed to 1,2- $C_{10}H_6ClNO$, HCl which was converted into 1,2- $C_{10}H_6ClNO$, according to the directions given by Hantzsch and Blagich (Ber. 33, 2541 (1900)) for the prepn. of 2- $C_{10}H_6NO$, from β - $C_{10}H_7NH_2$. The 1,2- $C_{10}H_6ClNO$, obtained in this way was identical with a specimen prepd. from 1,2- $C_{10}H_6(NH_2)NO$, by the Sandmeyer reaction (Hodgson and Kaiser, C. A. 20, 1074), it m. 70° . On heating with alc. NH_3 for 6 hrs. in a sealed tube at 180° , the $C_{10}H_6ClNO$ was converted into 1-amino-2-nitronaphthalene, m. 111° . On heating with alc. $MeNH_2$, 1-methylamino-2-nitronaphthalene, m. 114° , was formed and with alc. $EtNH_2$, 1-ethylamino-2-nitronaphthalene, m. 76° , the ammes reacting more easily than NH_3 , as is usually the case. On heating with alc. Na_2S , 2,2'-dinitrodinaphthyl disulfide, m. 204° , is obtained. C. F. VAN DUIN

Condensations of secondary amines with naphthols and aldehydes. H. WALLACH, R. ISONDI and JOSEPH B. LITTMAN. *J. Am. Chem. Soc.* 53, 1531-2 (1931). Cf. C. A. 24, 2155.—Naphthol (0.1 mol.) and 0.1 mol. of aldehyde in 10 cc. abs. EtOH were treated with 0.1 to 0.15 mol. of sec. amine and allowed to stand from 2 days to 5 months; the yields varied from 20 to 70%. Bz and Pr_2NH with β - $C_{10}H_7OH$ give 1-(α -dipropylaminobenzyl)-2-naphthol, m. 85° ; p - $MeOC_6H_4CHO$ and Me_2NH give 1-(α -dimethylamino- p -methoxybenzyl)-2-naphthol, m. 132° ; piperidine gives 1-(p -methoxy- α -1-piperidylbenzyl)-2-naphthol, m. 131° ; piperonal and Me_2NH give 1-(α -dimethylaminopiperonyl)-2-naphthol, m. 120° ; α - C_6H_5OH , Bz and piperidine give 2-(α -1-piperidylbenzyl)-1-naphthol, m. 110° . C. J. WHEAT

Optical activity and the polarity of substituent groups. XVII. 1-Menthyl hydrogen naphthalate, its alkali salts and methyl ester. H. GARDON RUEL and ANDREW McLEAN. *J. Chem. Soc.* 1931, 664-72, cf. C. A. 25, 249.—Menthyl (25-30 g.) in 50 cc. hot PhMe contg. 0.04 g. Na, added to a hot soln. of 8 g. naphthalic anhydride in 1000 cc. PhMe and heated 2 hrs. on a steam bath, gave 11 g. 1-menthyl 11-naphthalate (I), n. 141° (decompn., preheated bath) and then solidifies and m. 271° (m. p. of the anhydride), at room temp. the ester decomps. slowly. Na salt, very sparingly sol. in H_2O , also slowly decomps. at room temp.; K salt, rhombic crystals, is less stable than the Na salt, Li salt.—The Ag salt and MeI give the Me ester (II), in 96.5-7% (slight decompn.). In C_6H_6 , c. 4.004, II shows $[\alpha]_{D_{20}}^{20} = -89.1^{\circ}$, $[\alpha]_{D_{25}}^{25} = -120.1^{\circ}$, $[\alpha]_{D_{30}}^{30} = -117.3^{\circ}$, $[\alpha]_{D_{35}}^{35} = -208.2^{\circ}$; in $MeNO_2$, c. 3.093, the values are -35.3° , -48.3° , -59.1° , -133.8° ; in C_6H_6 , c. 4.014, I shows the values 117.1° , 165.0° , 201.3° and 397.0° ; in $MeCN$, c. 4.008, -35.2° , -47.4° , -55.8° , -82.3° . The influence of the MeO_2C substituent raises the rotatory power of the *peri* naphthoate, as well as of the α -lactonate, when the rotations are measured for the esters in EtOH soln. A remarkable fall in rotatory power occurs when the CO_2H group in I is modified by salt formation. The Li, Na and K salts all gave positive rotations under the conditions of the expt., the values not differing greatly from one another and being in the order $K < Na < Li < H$. In $MeOH$, c. 1, the following rotations of 1-menthyl 8-substituted 1-naphthoates are given: $[\alpha]_{D_{20}}^{20}$ CO_2Me , -34.1° ; Cl , -30.2° ; CO_2H , -218° ; CO_2Li , 0.67° ; CO_2Na , 4.25° ; CO_2K , 11.0° . C. J. WHEAT

Dyes derived from acenaphthenequinone. II. Azine and azonium derivatives. S. K. KUMAR GUPTA. *J. Chem. Soc.* 1931, 582-9, cf. Sircar and G., C. A. 18, 1202.—3-Bromoacenaphthenequinone, m. 218° , and α - $C_6H_5NH_2$ give 3-bromoacenaphthalphenazine, cream, m. 272° ; 3-Cl deriv., m. 276° , gives a reddish brown soln. in concd. H_2SO_4 ; 3,4-(H_2N) $C_{10}H_6$ gives 3-bromoacenaphthalolazine, cream, m. 270° ; 3-Cl deriv., cream, m. 250° ; H_2SO_4 gives a brownish red color. 2,3-Diaminophenazine gives 3-bromoacenaphthalphenazinazine, which does not melt and dyes wool light chocolate shades; 3-Cl deriv., gives greenish black solns. in concd. H_2SO_4 and dyes wool light chocolate shades from an acid soln. 1,2- $C_{10}H_6(NH_2)_2$ gives 3-bromoacenaphthalnaphthazine, lemon yellow, in 200%, and dyes wool in pleasant lemon-yellow shades; 3-Cl

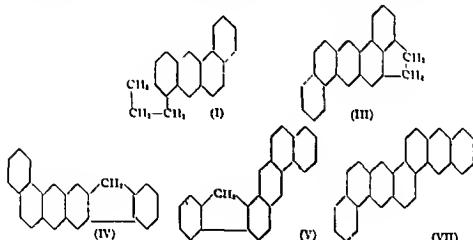
deriv, lemon yellow, m 274° , and dyes wool the same color. $1,2,5\text{-C}_{11}\text{H}_7(\text{NH})_3\text{SO}_3\text{H}$ gives 3 bromoacronaphthananaphthazine-3' sulfonic acid, canary yellow, 3-Cl deriv, brownish yellow, does not m 300° , concd H_2SO_4 gives blood red solns., wool is dyed light yellow shades. Phenyl amino- β -naphthylamine gives phenyl-3-chloroacronaphthananaphthazinone nitrate, yellow-orange, m above 300° (decompn), concd H_2SO_4 gives a violet color, wool is dyed yellow-orange shades from an acid bath. The tinctural properties of some of these compds. are not less developed than those of the available, most nearly corresponding phenanthraquinone derivs. C. J. WEST

The anomalous decomposition of the tetrazo derivative of 2,2'-diamino-1,1'-binaphthyl. A. CORBELLINI AND L. BARBARO. *Atti accad. Lincei* 12, 445-51 (1933) — It has already been shown that tetrazotized 2,2'-diamino-1,1'-binaphthyl decomps. giving an acidic deriv (I), m $250\text{-}2^{\circ}$ (C and Debenedetti, C A 24, 359), from H_2SO_4 or HCl solns. With HBr and Cu_2Cl_2 the normal di Br derivs are formed. Purified, I m $260\text{-}5^{\circ}$ and has the compn $\text{C}_{22}\text{H}_{12}\text{O}_2\text{N}_4$. With the Ag salt and MeI or EtI, the resp. esters, m 158° and 143° , are formed. The formation of a monobasic acid with 2 N atoms and the evolution of N_2 during decompn. indicate that one N N group behaves differently from the other. It is assumed that one of the N N groups in $(\text{C}_{22}\text{H}_{12}\text{N}_4\text{X})_n$ reacts normally to the —OH deriv, $\text{HO}(\text{C}_{11}\text{H}_7)_2\text{C}_2\text{H}_4\text{N}_2\text{X}$, N, while the other as a normal diazohydrate, $\text{HO}(\text{C}_{11}\text{H}_7)_2\text{C}_2\text{H}_4\text{N}_2(\text{O})\text{NH}$, oxidizes the double bond, then either of the tautomers $\text{HO}(\text{C}_{11}\text{H}_7)_2\text{CH}(\text{N})\text{CH}(\text{N})\text{N}(\text{C}_{11}\text{H}_7)_2$ or $\text{HO}(\text{C}_{11}\text{H}_7)_2\text{CH}(\text{N})\text{C}(\text{N})\text{N}(\text{C}_{11}\text{H}_7)_2$.

would result. The reactions of I show its resemblance to $\text{PhCH}(\text{COOH})_2$. On sublimation, it loses CO_2 , forming a neutral product m. 154° , its compn. being $\text{C}_{22}\text{H}_{12}\text{N}_4$. KMnO_4 oxidizes it, producing (1) an aldehyde $\text{C}_{22}\text{H}_{12}\text{O}_2\text{N}_4$, m 230° , (2) an acid which has not been purified but is probably related to benzoic or phenylglyoxylic acid, and (3) an acid, m 287° . On bromination of the esters unstable derivs. are obtained which are apparently di Br addn. products. A. W. COWLEY

Polycyclic aromatic hydrocarbons. II. The non-existence of 1,2,7,8-dibenzanthracene. JAMES W. COOK. *J. Chem. Soc.* 1931, 457-9, cf. C A 24, 4020 — By pyrolysis of 2 methyl 1,1' and 1,2'-dinaphthyl ketones. Clar (C A 23, 2844) and Fieser and Dietz (C A 23, 5472) obtained compds. which they believed to be, resp. 1,2,7,8- (I) and 1,2,5,6-dibenzanthracenes (II). A further study of these compds. shows that they have the same fluorescence spectrum (Heeger, C A 24, 5369) and yield the same derivs. That the compd. is II is shown by its synthesis by Weitenbeck and Klinger (C A 13, 423). Homer (C A 5, 1076) claimed that I was formed by the action of AlCl_3 on $\text{C}_{22}\text{H}_{14}$, $\text{C}_{22}\text{H}_{14}\text{Cl}_2$ repetition of the work showed that the product is perylene. It appears, therefore, that I has not yet been prepd. III. Derivatives of 1,2,5,6-dibenzanthracene. Ibid. 489-99 — The following compds. were prepd. in a further effort to identify the carcinogenic principle in coal tar: 2,6- $\text{C}_{22}\text{H}_{14}\text{Me}_2$ and α - $\text{C}_{22}\text{H}_{14}\text{COCl}$ with AlCl_3 in CS_2 give 50% of 2,6-dimethyl 1,1'-dinaphthyl ketone (I) yellow, m $162\text{-}3^{\circ}$, 2,7-di Me deriv (II), m $161\text{-}2^{\circ}$ (70% yield), 2,7- $\text{C}_{22}\text{H}_{14}\text{Me}_2$ and β - $\text{C}_{22}\text{H}_{14}\text{COCl}$ give 40% of 2,7-dimethyl-1,2'-dinaphthyl ketone (III), m $113\text{-}5\text{-}4\text{-}5^{\circ}$, 1,6- $\text{C}_{22}\text{H}_{14}\text{Me}_2$ and R_2Cl give β -benzyl 1,6-dimethylnaphthalene (IV) m $77\text{-}8^{\circ}$, it is scarcely attacked by alk. KMnO_4 after boiling 10 hrs. 2- $\text{C}_{22}\text{H}_{14}\text{Me}$ and 4- $\text{MeC}_{22}\text{H}_{14}\text{COCl}$ give a poor yield of 2,4'-dimethyl 1,1'-dinaphthyl ketone (V), m $120\text{-}1^{\circ}$, 2,7,4'-tri Me deriv (VI), m 140° . Heating V at $440\text{-}50^{\circ}$ until H_2O was no longer eliminated gives 1,2,5,6-dibenzanthracene (VII), I under the same conditions gives the 3' Me deriv of VII, which was also obtained from the crude condensation product of 2,6- $\text{C}_{22}\text{H}_{14}\text{Me}_2$ and 4- $\text{MeC}_{22}\text{H}_{14}\text{COCl}$. III, on pyrolysis, gives the 2' Me deriv of VII, m $256\text{-}7\text{-}5^{\circ}$, as does the pyrolysis of II, of the ketone from 1,6- $\text{C}_{22}\text{H}_{14}\text{Me}_2$ and 2- $\text{MeC}_{22}\text{H}_{14}\text{COCl}$ and of VI. In each case the Me group in the α -position to a condensed ring is eliminated. The structures of the ketones were not rigidly established but follow from general considerations: 1,2,5,6-Dibenzanthraquinone (10.5 g) and MeMgI in C_6H_6 give 10.2 g. of 9,10-dihydroxy-9,10-dimethyl 9,10-dihydro-1,2,5,6-dibenzanthracene (VIII), m $245\text{-}50^{\circ}$, concd H_2SO_4 gives an intense concolor blue color, reduction with Zn and AcOH gives only reagent substances but HI in glacial AcOH gives 9,10-dimethyl 1,2,5,6-dibenzanthraquinone, yellow, m $205\text{-}6\text{-}5^{\circ}$, and gives an intense violet soln. in concd H_2SO_4 , the same compd. was obtained from the so-called 1,2,7,8-dibenzanthraquinone of Clar. Reduction of VIII with red P and HI in glacial acetic acid gives a mixt. of the *cis* and *trans* 9,10-dimethyl 9,10-dihydro derivs, m $207\text{-}9^{\circ}$ and $277\text{-}8^{\circ}$ resp. the 2 were sep'd by crystn. from $\text{C}_{11}\text{H}_7\text{Me}$, the *cis* form being the more sol. the *cis* form shows a feeble fluorescence in ultra violet light. The 9,10-di Bu deriv corresponding to VIII, m $214\text{-}5^{\circ}$, and gives the same color in H_2SO_4 , a Bu ether (I'), m $225\text{-}6^{\circ}$, was obtained when HCl was used to decomp. the Grignard reagent, reduction of the diol with Zn and AcOH gives the 9,10-di Bu deriv of

VII, yellow, m 143.5–4.5°, warm concd H_2SO_4 gives a purple soln with a dark red fluorescence. The 9,10-dibenzyl deriv of VII, m 195–201° (decompn), purified through the red picrate, was obtained by reduction of the corresponding *diol*, m 249–51°, which gives an orange-red soln in concd H_2SO_4 . Reduction of VII with Na and Al(OH)_3 gives an octahydro deriv, m 188.0°, also obtained from Clir's compd. A preliminary note is added on the isolation of a hydrocarbon, m 153.0°, from the high boiling constituents of coal tar, which, on oxidation, gives a mixt of quinones. IV. Condensed derivatives of 1,2-benzanthracene. *Ibid* 499–507. This work was for the purpose of prepg. an unsubstituted hydrocarbon having the same fluorescent spectrum as that of the powerful cancer producing mixts of unknown composition. This was largely realized in the compds III and IV or V. 2 $\text{MeC}_{10}\text{H}_7\text{COCl}$ and Ph_2 give only 1,2-(*p*- $\text{PhC}_{10}\text{H}_7\text{CO}$) $\text{C}_{10}\text{H}_7\text{Me}$ (*m* $\text{MeC}_{10}\text{H}_7$), and $\alpha\text{-C}_{10}\text{H}_7\text{COCl}$ with AlCl_3 give 4,4'-di- α -naphthyl-3,3'-dimethylbiphenyl, cream, m 159.5 (lit 5), pyrolysis gives 1,2,1',2'-dibenz-6,6' (or 7,7')-bunthryl, m and decompn above 310°, sublimes about 375° at 2 mm. It may be a mixt of the 2 isomers. The ketone from hydriindene and 2 $\text{MeC}_{10}\text{H}_7\text{COCl}$ could not be crystal. pyrolysis gives cyclopenteno 1,2-benzanthracene (probably I), m 199–200°, concd H_2SO_4 gives a crimson soln. This was also obtained by pyrolysis of the crude ketone from hydriindene 5-carboxyethyl chloride and 2 $\text{MeC}_{10}\text{H}_7\text{MgI}^2$. 2 Methyl-5',6',7',8'-tetrahydro-1,2'-dinaphthyl ketone m 122.5–3.5, results in 70% yield from tetraol and 2 $\text{MeC}_{10}\text{H}_7\text{COCl}$, dehydrogenation with δ at 210–31° gave a resin from which 2 methyl 1,2'-dinaphthyl ketone was isolated. pyrolysis gave only 1,2,5-bibenzanthracene (II). Acenaphthene and 2 $\text{MeC}_{10}\text{H}_7\text{COCl}$ give 15% of 7-(2'-methyl 1'-naphthyl)acenaphthene, m 184.5°, pyrolysis gives phenanthracenaphthene (III) golden yellow m 231–2°, sublimes at 330° and 3–4 mm. It was purified through the chocolate-brown picrate, concd H_2SO_4 gives a fugitive bluish violet color. distn gives II, the 5 membered ring being eliminated. Fluorene gives 85% of 2-(2'-methyl 1'-naphthyl)fluorene, m 169–70°, pyrolysis gives 2,3-phenanthra 3',2'-fluorene or 2,3-phenanthra 1',2'-fluorene (IV or V), pale yellow, m 302–4°, concd H_2SO_4 gives a pale blue soln with a pinkish red fluorescence, which becomes emerald green on standing or gentle warming. 3-Phenanthrolyl chloride, m 110°, from the acid and SOCl_2 , gives with $\alpha\text{-MeC}_{10}\text{H}_7\text{MgI}^2$ a ketone, which could not be crystal but gave on pyrolysis 2',3'-naphthyl 2,3-phenanthrene, m 261–4°, with 2 $\text{MeC}_{10}\text{H}_7\text{MgI}^2$ there results 3-(2'-methyl-1'-naphthyl)phenanthrene, m, 145–6°, giving on pyrolysis 2',3'-phenanthra-2,3-phenanthrene (VI), canary-yellow, m 311–12°, giving an indigo-blue color in concd H_2SO_4 changing to emerald-green on warming, a small quantity of another hydrocarbon, orange, m 245–8°, is formed, probably isomeric with VI. 2 $\text{MeC}_{10}\text{H}_7\text{MgI}^2$ and 1-anthryl chloride give a ketone as a viscous oil, which, on pyrolysis, gives 2',3'-phenanthra 1,2-anthracene, canary-yellow, m 261–2° (decompn). 1-Benzoyl-5 (1'-naphthyl)-2,6-dimethylnaphthalene, m 206–8°, in 80 g yield from 10 g 2,6-dimethyl 1,1'-dinaphthyl ketone, the 2'-naphthyl deriv, m 222–3°, pyrolysis of the latter gives 4,5-benz-10,11-(1',2'-naphthyl)chrysene (VII), golden yellow, m 435–40° (decompn). Attempts to prep diketones by condensing 2,7-dimethyl 1,1'- and 1,2'-dinaphthyl ketone with BrCl were fruitless.



C. J. West

Syntheses of antiseptic derivatives of indan-1,3-dione. II. Interaction of alkyl-malonyl chlorides with *p*-tolyl methyl ether. THOMAS K. WALKER, ARTHUR J. SUTHERS, LESLIE I. ROF AND HERBERT SHAW. *J. Chem. Soc.* 1931, 514-20, cf. *C. A.* 25, 2140—*p*-MeC₆H₄OMe (5 g) and 58 g CH₂(COCl)₂ in 50 cc PhNO treated with 11.5 g AlCl₃ during 30 min and the mixt. warmed at 61° for 45 min, gave 50% (4 g) of 4-hydroxy-7-methylindan-1,3-dione (I) lemon yellow, m 258°. *p*-MeC₆H₄OH (24 g) and 34 g MeCH(COCl)₂ gave 0.8 g of the 2,7-di Me deriv (II), m 233°. *p*-MeC₆H₄OMe and EtCH(COCl)₂ gave 20% of the 2 Et deriv (III), m 167°. EtC(COCl)₂ and *p*-MeC₆H₄OMe gave the 2,2-di Et deriv (XII), m 199-200°, if the reaction was carried out in CS₂ the main product was the corresponding 4-MeO deriv, oily, heating with AlCl₃ gives a good yield of III. EtCH(COCl)₂ and *p*-MeC₆H₄OMe give 4-hydroxy-7-methyl-2-propylindan-1,3-dione (IV) m 187°, this was also obtained in quant yield from *p*-MeC₆H₄OAc. BuCH(COCl)₂ gives the 2 Bu deriv (V), m 165°. The same compd. was also obtained with *p*-MeC₆H₄OH with AlCl₃ or FeCl₃ and from PCl₅. *p*-MeC₆H₄OAc. BuCH(COCl)₂ and PhNO, the product finally being treated with AlCl₃ 2-1m deriv (VI) m 144° (50% yield) 2-hexyl deriv (VII), m 135° (31% yield); 2-heptyl deriv (VIII) m 121° (50% yield) 2-iso Pr deriv (IX) m 224° (16% yield) 2-iso Bu deriv (X) m 152.5° (47% yield) 2-iso Am deriv (XI), m 142° (44% yield). The following figures were obtained for the equimol. 1%OH coeff. of bacteriostatic power: I 64 II 30 III 37 IV 64 V 170 VI 440, VII 1305 VIII 1170, IX 106 X 133, XI 360 XII 67. The concns of V required to inhibit *S. proteogenes albus*, *B. subtilis* and *B. pili* for 48 hrs at 37° were 1:25,000 1:25,000 and 1:3,000, resp. the same organ. was inhibited under the same conditions by VIII in resp. concns. of 1:500,000, 1:333,000 and 1:500,000. The attachment of 1 of the lower alkyl groups in place of a H atom, to the C atom in the medial position with respect to the 2 CO groups results in a diminution of antiseptic potency. As the length of the alkyl chain is increased, this effect becomes neutralized, groups higher than Pr giving rise to considerably enhanced activities, the progressive increase of which appears to have reached its limit with the introduction of the C₁₂H₂₅ radical.

C. J. WEST

Anthraquinone azo compounds. II. Insoluble disazo dyes from 1,5-diamino-anthraquinone. TOSIO MAKI. *J. Soc. Chem. Ind., Japan* 34, Suppl. binding 51.6 (1931), cf. *C. A.* 24, 4780—Tetraazotized 1,5-diaminoanthraquinone when coupled gives dyes as follows: with a C₁₂H₁₁NH₂ violet brown, m 230°, with a C₁₁H₁₁NH₂ violet brown m 269.9°, PhNH₂ orange, m 243.5°, m C₆H₅(NH₂)₂ dark brown, m 259.6°, a C₁₀H₇OH violet brown m over 300°, PhOH, yellow brown, m 280.8°, m-C₆H₄(OH)₂ red brown m 400°.

V. I. HARRINGTON

Velocity measurements on the opening of the furan ring in hydroxymethylfurfuraldehyde. II. H. P. TRUNISSEN. *Rec. trav. chim.* 50, 120 (1931) cf. *C. A.* 24, 4782—In the 1st paper measurements of the hydrolysis of hydroxymethylfurfuraldehyde under various conditions were described, water being the solvent, the present paper deals with the same subject. mixts. of water and MeOH or EtOH being used as a solvent. The 1st series of measurements was carried out with EtOH water mixts. at the *b. p.* of the mixt., the following results being obtained with 0.5 N HCl as a catalyst: 5% EtOH, *k* = 0.00129, 10% 0.00106, 15% 0.00083, 25% (b 90.9-91.2°), 0.00046, 30% (b 89.9-90.0°), 0.00034, 40% (b 87.6-87.7°), 0.00023, 50%, 0.000125, 70%, 0.00003. Although these reaction consts. do not refer to the same temp., they show clearly that the velocity of the hydrolysis of hydroxymethylfurfuraldehyde into formic and levulinic acids decreases on increasing the EtOH content of the mixt. Moreover, it was observed that the reaction const. always decreases during an expt., most probably alcoholysis taking place simultaneously, in mixts. with the higher EtOH contents the quantity of humus formed is much less than in water and in mixts. with the lower EtOH content. The same results were obtained on carrying out these measurements in a sealed tube or a closed flask at 100° again a continuous decrease of the reaction const. being observed during the course of each expt. With 0.5 N HCl as a catalyst, the following results were obtained at 100°: 10% EtOH *k* = 0.00150, 15% 0.00131, 25% 0.00090, 35%, 0.00055, 50% 0.00022, 75%, 0. From these results the temp. coeff. of the reaction is calculated to be about 2.0. With mixts. of MeOH and water again the same phenomena were observed as regards both the continuous decrease of the reaction const. during each expt. and the retarding effect of the alc. in general. With 0.5 N HCl as a catalyst the following figures were obtained at 100°: 10% MeOH, *k* = 0.00111, 20%, 0.00063, 30%, 0.00020, 40% about 0. The general retarding effect of the addn. of MeOH or EtOH is in accordance with the researches of Wechurzen who recommends the use of a satd. soln. of HCl in EtOH instead of concd. HCl in both the Selivanov and the Bauduin reactions (*C. A.* 12, 1004, 2029) in order to prevent decompn. of the hydroxymethyl-

mahogany wood (West Indies) 6 g I and 6 g III; 8 kg Australian kino (IV) from *Eucalyptus terminalis* 16 g I and 22 g III; 6 kg IV from *E. leucocylon*, 5 g I and 10 g III; 9 kg IV from *Angophora intermedia* 14 g I and 8 g III; 86 kg heartwood of *Anacardium occidentale*, 21 g II; 4 kg Malabar kino (V) from *Myristica malabarica*, 3 g I acetatechol (VI) and 11 g III; 5 kg V from *Pterocarpus marsupium*, 9 g VI and 17 g III, 6 kg IV from *E. corimbosa*, 4 g VI and 9 g III; 28 kg *Uncaria gambir* leaves, 50 g d gambir catechol (VII) and 6 g d isomer (VIII), 28 kg *U. acida* leaves, 17 g VII and 3 g VIII, 11 kg Chinese rhubarb, 17 g VII and 3 g VIII; 2 kg Guarana paste, 5 g VII and 8 g VIII, 10 kg kola nuts, 58 g VII and 8 g VIII. C. J. West

Polyhalogenated ketones of indole. G. SANNA with FARRAELLA MASSIDA. *Gazz. chim. ital.* 61, 60-74 (1931). —The present paper describes the prepn of 2 polyhalogenated ketone derivs of methylketole and the behavior of these derivs with alkalis and with NH_4OH which differed greatly in the 2 cases. CH_3COCCl in anhyd. Et_2O added dropwise to magnesium methylindole (I) (1 mol), keeping cold, kept ice-cold several hrs., ice water added, the Et_2O layer sep'd, the aq. layer filtered, the residue washed with water allowed to dry, dissolved in hot MeOH , cooled and the purification repeated several times, yields almost 100% of β -dichloroacetyl α methylindole (α methylindacetyl chloride) (II) m 195° gives green solns in cold concd. H_2SO_4 and solns in hot concd. HNO_3 which are at first green and then yellow. It is almost insol in boiling dil acids. It is slightly sol in hot aq. NaOH and KOH and these yellow solns give amorphous red ppts when acids are added. It is lat more sol in alc. KOH , and addn of water or of dil acids to these solns ppts rose-colored amorphous products. It is insol in hot NH_4OH , but sol in hot alc. NH_3 . Its solns in hot $\text{C}_6\text{H}_5\text{N}$ do not ppt on cooling, but on addn of dil LiOH a ppt contg Cl is formed (perhaps II). It reduces hot NH_3 , AgNO_3 and slowly reduces Fehling soln. It gives ppts of various colors with the general reagents of aldehydes. The Cl of II is very stable, and there is no decompn when a 1% aq. soln is boiled for several hrs. The Clean, however, be removed by boiling several hrs in concd. alc. KOH . In this latter case the product is amorphous and is difficult to purify, but is probably β -diacetoacetyl α methylindole. Fusion with KOH , carried out in the same way as with other monohalogenated derivs, yields α methyl β indolecarboxylic acid. Oxidation is more difficult than with indacyl chloride (III), and it is necessary to boil it for a long time in more concd. KMnO_4 , the products are, however, the same (cf C A 24, 2127) from II and from III. The ease with which II reacts with KBr and KI suggested the same reaction with II. II in MeOH and concd. aq. KBr (2 mols), refluxed 48 hrs., cooled, filtered, the residue washed with water, dried, fractionally crystd from MeOH (the greater part is II), gives a small yield of a compd which is probably

α methylindacetyl dibromide, $\text{C}_8\text{H}_7\text{NH CMe CCOCHBr}_2$, m 178°. Following the same procedure with KI the yield of the compd, which is probably α methylindacetyl diiodide, is still smaller. This difficulty of substitution of Cl by other halogens is a new example of the great stability of Cl in dihalogenated derivs. II (2 g) and 5% aq. KOH (2 mols), refluxed 12 hrs (60% of II is still unaltered), filtered, the filtrate acidified with dil HCl extd with Li_2O , the ext. dried with anhyd. Na_2SO_4 , concd on a water bath, the yellow oil kept *in vacuo* until crystd and recrystd from Et_2O , yields α methyl β indyl

glycolic acid $\text{C}_8\text{H}_7\text{NH CMe CCH(OH)CO}_2\text{H}$ (IV), m 90° (decomps to CO_2 and α methylketoylcarbinol), insol in dil acids, its soln in concd. H_2SO_4 is at first yellow but this changes to red, in concd. HNO_3 it is rose-yellow, in hot concd. HCl (rose, in NH_4OH straw yellow (addn of HCl ppts a rose-colored amorphous powder), and in caustic alkalis yellow (repp'd in cryst. form by dil acids), it reduces Fehling soln immediately and NH_3 , AgNO_3 without formation of a metal mirror, does not give any color with Schiff reagent. Aq. BaCl_2 added to IV in aq. KOH (calcd quantity) ppts the Ba salt of IV, could not be fused. Dil aq. AgNO_3 added to IV (calcd quantity) in NH_4OH ppts the Ag salt of IV, turns a rose color in light m 247°. CH_3COCCl in anhyd. Et_2O , added dropwise to 1 (1 mol) keeping cold let stand 2 hrs at 0°, ice water added, the Et_2O layer sep'd the aq. layer extd with Et_2O , the combined Et_2O extd dried with anhyd. Na_2SO_4 , filtered the Et_2O evapd, the residual oil washed with petr. ether, allowed to crystallize and the product purified by recrystn successively from boiling C_6H_6 and MeOH (with animal charcoal) yields almost 100% of α methyl β trichloro

acetylindole (methylketoylchloroform) $\text{C}_8\text{H}_7\text{NH CMe CCOCCl}_3$ (V), m 167°, barely sol in hot dil acids and in hot concd. HCl its soln in hot concd. HNO_3 is at first yellow, then rose, in hot caustic alkalis rose (odor of CHCl_3) and in $\text{C}_6\text{H}_5\text{N}$ also rose (addn of dil EtOH ppts V again), it is insol in NH_4OH reduces hot Fehling soln (this is prob-

ably due to CHCl_3 formed by the action of alkali on the reagent, which transforms it into formate), reduces NH_4AgNO_2 soln V (1 g) and 5% aq KOH (3 mols), refluxed 6 hrs (most of the V is still undissolved), filtered, the filtrate cooled, acidified with dil H_2SO_4 , extd with Et_2O , the ext dried with anhyd Na_2SO_4 , coned to a small vol, the residual soln evapd *in vacuo* in a desiccator yields a compd (VI), red, m approx 80° , insol in dil acids, slightly sol in hot coned HCl , its soln in coned H_2SO_4 is at first yellow but turns red, it is partially sol in NH_4OH (yellow color), slightly sol in cold caustic alkalis, easily sol in hot caustic alkalis and the yellow soln thus formed becomes rose colored on addn of HCl and a rose colored amorphous ppt. is formed. VI reduces Fehling soln and NH_4AgNO_2 soln (without formation of a mirror). It does not contain Cl , behaves as a weak acid and the closeness of its m p to that of IV suggests that IV and VI are the same compd, but this was not ascertained. V (1 g) and dil. alc KOH (3 mols), heated gently, dild with water (CHCl_3 is evolved), acidified with HCl , extd with Et_2O , the ext dried, coned and let stand *in vacuo*, deposits α -methyl β

indolecarboxylic acid, $\text{C}_{11}\text{H}_9\text{NH CMe CCO}_2\text{H}$, m 174° (decompn). The same reaction occurs with cold coned alc KOH and with hot coned aq KOH . A suspension of V (1 g) in 30% aq K_2CO_3 , refluxed several hrs., filtered, the filtrate cooled, the ppt filtered, dried and recrystd from EtOH , yields methylketoylformic acid (α -methyl β -indylglyoxylic

acid), $\text{C}_{11}\text{H}_9\text{NH CMe CCOC}_2\text{H}_4$ (VII), golden yellow, m 186° (decompn. to CO_2 and α -methyl β -indolic aldehyde), insol in dil acids, its soln in coned H_2SO_4 is yellow, it is sol in caustic alkalis and in NH_4OH , does not reduce Fehling soln or NH_4AgNO_2 . Its formation is therefore as follows $2\text{V} + 3\text{K}_2\text{CO}_3 + \text{H}_2\text{O} \rightarrow \text{VII} + 3\text{CO}_2 + 6\text{KCl}$. As described for II, attempts were made to substitute the Cl of V with Br and I , to which end alc. V was treated with KBr (3 mols), but there was no reaction even after boiling for several hrs. Further expts to obtain the Br and I derivs are in progress. II (2 g), suspended in coned NH_4OH , heated in a sealed tube for 6 hrs at 100° , cooled, the ppt dissolved in boiling water and cooled, ppts 100% of α -methyl β -indolecarboxylic acid

amide, $\text{C}_{11}\text{H}_9\text{NH CMe CCONH}_2$ (VIII), silky, m 218° , its solns in boiling dil acids are yellow-rose, and in coned H_2SO_4 yellow, it is insol in cold and only slightly sol in hot aq alkalis. The formation of VIII should be thus $\text{II} + \text{NH}_3 \rightarrow \text{VIII} + \text{CH}_2\text{Cl}_2$, and since CH_2Cl_2 was not detected in the reaction product it is probably transformed first into HCHO and thence into hexamethylenetetramine. Further expts on this reaction are in progress. Under the same conditions, the reaction of V and NH_3 did not give definite results. The V hardly dissolved even after prolonged heating, and the soln, which turned yellow, deposited, after long standing, a red powder, the quantity of which was too small for analysis.

C. C. DAVIS

Preparation of isatin from isonitrosoacetanilide according to Sandmeyer's method. J. P. WIBAUT AND M. C. GEERLING. *Rec trav chim* 50, 41-3 (1931).—On prep of isatin by heating isonitrosoacetanilide and coned H_2SO_4 to 70° and pouring the reaction mixt on ice (Sandmeyer, *C A* 13, 1840; Marvel and Hiers, *C A* 20, 193) it is often observed that the isatin seps from the soln only after inoculation or after scratching with a glass rod, while in other cases isatin monoxime, m $201-2^\circ$, is obtained instead of isatin, the oxime compd may also be obtained from solns from which isatin itself has already crystd. The NH_4OH , necessary for the formation of the isatin monoxime, is produced by hydrolysis of the isonitrosoacetanilide, which is not converted into isatin in the dil H_2SO_4 ; it was shown that this hydrolysis does not take place in the coned H_2SO_4 .

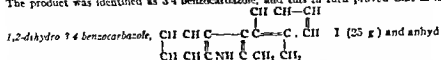
C. F. VAN DUIN

Action of hydrogen peroxide on benzo-6,7-diketo-2,3-dihydro-2,3-dithionaphthene in sodium hydroxide solution. R. STOLLÉ AND W. BADSTÜBER. *J. prakt Chem* 129, 309-11 (1931).—The action of H_2O_2 on benzo-6,7-diketo-2,3-dihydro-2,3-dithionaphthene in NaOH gives bis(2-carboxy-1-naphthyl) disulfide, m 156° . KMnO_4 is reduced in the cold. Titration with 0.1 N alkali to neutrality gives the di-K salt; if excess KOH is added the acid is apparently decompd into 1,2- $\text{C}_{10}\text{H}_7(\text{SO}_3\text{K})\text{CO}_2\text{K}$ and 1,2- $\text{C}_{10}\text{H}_7(\text{SK})\text{CO}_2\text{K}$.

C. J. WEST

Carbazolic derivatives. E. GHIGI. *Gazz. chim. ital* 61, 43-6 (1931).—In previous expts (cf. *C. A.* 24, 3797) condensation of β tetralone phenylhydrazones (I) with dil H_2SO_4 led to the formation of a compd (II), m. 100° , which may have had either one of two formulas, for a choice between which there were at the time insufficient data. In the present paper, the identity of this compd. is definitely established. It was dehydrogenated by the method of Borsche (*C. A.* 2, 1716), *s. e.*, by heating II (5 g) mixed with PhO (equal vol) in a current of CO_2 and purifying the distillate by crystn from ligroin

The product was identified as 3,4-benzocarbazole, and this in turn proved that II is



1,2-dihydro-3,4-benzocarbazole, $\text{C}_{12}\text{H}_{11}\text{N}$ (25 g) and anhyd ZnCl_2 (83 g) in abs EtOH refluxed for 10 hrs, 5% aq HCl added, extd with Et₂O, the ext dried with anhyd Na₂SO₄, the Et₂O eliminated, dried *in vacuo* and the residue recrystd from C₆H₆, yield the compd C₁₂H₁₁N (III), m 224°, heated with anhyd (CO₂H)₂, it gives the blue color characteristic of carbazole derivs. III (1 g) and picric acid (1 g) in C₆H₆, refluxed until clear, cooled and the ppt recrystd from C₆H₆, yield the picrate C₁₂H₁₁N·C₆H₃(O₂H)₃ red m 146°. III, Ac₂O (4 parts by wt) and anhyd ZnCl₂ refluxed 15 min, cooled and the ppt recrystd from EtOH, yield the Ac deriv, m 289-90°. Following the method of Borsche as before (*loc cit*), but purifying the product from C₆H₆, 1,2-benzo-3,4-dihydrocarbazole (cf C A 24, 3797) and PbO form III. III is therefore 1,2-benzocarbazole, obtained by Borsche (*loc cit*) by oxidation of 1,2-benzo-3,4-tetrahydrocarbazole. C C DAVIS

Synthesis of 4-phenylthiazole-2-methanol and some of its derivatives. VIII. JOHN I CHIN AND TREAT B JOHNSON. *J Am Chem Soc* 53, 1470-3 (1931) — Br·OCH₂CSN₂ (40 g) 38 g BrCH₂Br and 20 cc C₆H₅N in 100 cc EtOH, heated 1 hr on the H₂O bath, give 40% of the benzate, m 73-4°, of 4-phenylthiazole-2-methyl alc (I), PhC·CH₂SC(CH₂OH)N, b₁₀ 210-3°, m 88-9°. Similarly, Br·OCH₂CSN₂ gives the

Et ether of I, b₁₀ 187-8°, unchanged by 40% HBr for 4 hrs, heating with 80% H₂SO₄ at 145-50° for 0.5 hr gives 65% of I. The acetate of I, b₁₀ 193°, m 40°. Bubbling HBr through I in Ac₂O while warming on the H₂O bath gives 90% of the bromide, b₁₀ 195°, the chloride, b₁₀ 194°. Oxidation of I with Na₂Cr₂O₇ in AcOH gives 4-phenylthiazole-2-aldehyde, b₁₀ 160-2°, phenylhydrazone, canary yellow, m 131-2°, KCN in EtOH gives 4-phenylthiazole-2-aldoxime, yellow, m 27-8°. C J WEST

Synthesis of 4-phenyl-2-acetylthiazole. IX. JOHN F OLIN AND TREAT B JOHNSON. *J Am Chem Soc* 53, 1473-5 (1931) — Benzoyl lacte thioamide (53 g) and 50 g BrCH₂Br in 150 cc EtOH, heated 5 hrs on a steam bath, give 35 g of the benzate, b₁₀ 212-4°, of 4-phenylthiazole-2-ethanol, PhC·CH₂SC(CH₂OH)Me, N, yellow, b₁₀ 191-

4°, m 76°. This could not be dehydrated by P₂O₅. Oxidation of the alc with Na₂Cr₂O₇ in AcOH gives a nearly quant. yield of 4-phenyl-2-acetylthiazole PhC·CH₂SCAc, N,

yellow, m 78-9°, phenylhydrazone, yellow, m 208-9°. Br in CCl₄ gives the 2-bromoacetyl deriv, buff, m 100-7°, when warmed in dil alk. soln or with an amine, it readily lost Br, indicating that the halogen occupies a reactive position in the Me radical of the ketone group. C J WEST

Synthesis of some new thiazole amines containing the catechol group. X. JOHN F OLIN AND TREAT B JOHNSON. *J Am Chem Soc* 53, 1475-7 (1931) — ClCH₂CH₂·CH₂NH₂ is converted into the phthalimido deriv, m 81°, in 80% yields with EtOH·(NH₄)₂S, after standing 72 hrs, there results γ-phthalimidobutyrothioamide, m 181-2°, with chloroacetocatechol this gives 4-(3,4-dihydroxyphenyl)-2-γ-phthalimidopropylthiazole, yellow, m 114-5°, heating with NaH, H₂O gives 4-(3,4-dihydroxyphenyl)-2-γ-aminopropylthiazole, isolated as the sulfate, C₁₂H₁₀O₄N₂S·H₂SO₄, m 226-8°, it crystallizes with 3.5 mols H₂O, m 120-40° (decompn). Phthalimidopropionitrile, m 147-8° (the yield is small owing to the formation of acrylic nitrile) phthalimidopropionic thioamide, m 187-9°, 4-(3,4-dihydroxyphenyl)-2-β-phthalimidoethylthiazole, m 203-5°, NaH, H₂O gives 4-(3,4-dihydroxyphenyl)-2-β-aminoethylthiazole, whose sulfate crystallizes as needles. The pharmacol study is in progress. C J WEST

Anthracene derivatives V. 2,1-Thionidoids of the anthraquinone series. PAUL RUGGIE AND WALTER HEITZ. *Helv Chim Acta* 14, 257-70 (1931), cf C A 25, 947 — The starting material for these syntheses was 1-amino-2-mercaptoanthraquinone (I) (cf Ger pat 290,054, Friedl 12, 439 (1916)), prepd. by adding rapidly with stirring 40 g 1-aminoanthraquinone to 800 g technical crystd Na₂S previously heated in a paraffin bath and allowed to cool to 70° the temp was increased to 140° during 4 hrs and kept at 140° for 6 hrs the mixt was then poured into a porcelain dish, leached with 2 l of boiling H₂O, filtered through cotton to remove unchanged aminoanthraquinone and Na₂SO₄, the filtrate evapd to 0.5 vol., allowed to crystallize by standing overnight, it yielded 36-41 g I. When 3 g I was dissolved in 150 cc hot H₂O, then cooled to 40°,

12 cc 30% NaOH added, then 9 cc Me_2SO_4 added slowly with shaking, allowed to stand 0.75 hr at 70° , an equal vol of H_2O added, warmed 15 min on the water bath, filtered, the ppt washed with H_2O and EtOH , it gave 2.2 g of 1-amino-2-(methylthio)anthraquinone, red needles, m 186° . Treatment of 3 g I in 30 cc warm H_2O with BrCH_2COPh in 10 cc EtOH gave a ppt of 3.6 g 1-aminoanthraquinone 2-(thiomethyl) phenyl ketone (II), $\text{C}_{18}\text{H}_{13}\text{O}_2(\text{SCH}_2\text{COPh})(\text{NH}_2)$ red needles decompose with ring closure at 130° , the same ring closure of II with loss of H_2O to form 1-aminoanthraquinone-2-thiomethyl phenyl ketimide, $\text{C}_{18}\text{H}_{11}\text{O}_2\text{SCH}_2\text{CPhN}$, occurs when II is heated with AcOH , PhNO_2 , or

on dissolving in concd H_2SO_4 and pptg with H_2O , it forms blue black crystals, m 262° . A soln of 10 g of the Na salt of I in 120 cc H_2O at 80° was treated with 12 g $\text{ClCH}_2\text{CO}_2\text{H}$ (neutralized with Na_2CO_3) in 40 cc H_2O , allowed to stand 1.5 hr on the water bath, filtered while hot and the filtrate on the addn of NaCl gave 10.5 g of Na 1-aminoanthraquinone 2-thioglycolate (III), 4 g III on boiling with 20 cc Ac_2O gave 3.5 g 1-aminoanthraquinone-2-thioglycolic lactam (IV), $\text{C}_{18}\text{H}_{13}\text{O}_2\text{SCH}_2\text{CONH}$, IV is also formed from

III on heating with concd H_2SO_4 or on boiling with AcOH . Et 1-aminoanthraquinone 2-thioglycolate (V), prepd in 75% yield by adding $\text{BrCH}_2\text{CO}_2\text{Et}$ to 1,2 $\text{C}_6\text{H}_4(\text{CO})_2\text{C}_6\text{H}_3(\text{NH}_2)\text{SNa}$, red m 116.5° , when warmed with concd H_2SO_4 it forms the lactam IV. Diazotization of V (20 g) in 300 cc cold, concd H_2SO_4 with 6 g NaNO_2 in 40 cc H_2SO_4 with continued stirring 2 hrs after all of the NaNO_2 was added, then pouring into 1 l ice water, rapidly filtering, treating the filtrate with a cold, satd KClO_4 soln, filtering off the pptd diazonium perchlorate (which is remarkably stable) suspending the latter in H_2O , adding a KCN CuCN soln warmed to 70° (from 24.5 g CuSO_4 in 200 cc H_2O and 27 g KCN in 100 cc H_2O), heating 2 hrs on the water bath, filtering while hot, washing successively with boiling H_2O , dil HNO_3 and boiling H_2O , gave 16.4–18.6 g of Et 1-cyanoanthraquinone 2-thioglycolate, brown, m 222° , the corresponding Me ester similarly prepd, m 217° hydrolysis of these esters by boiling with a mixt of concd HCl and AcOH gave an almost quant yield of 1-cyanoanthraquinone 2-thioglycolic acid (VI), gray brown, m $260\text{--}70^\circ$ (decompn), and on boiling with $\text{C}_6\text{H}_5\text{N}$ forms the pyridine salt. Hydrolysis of 23 g VI by boiling with a mixt. of 184 cc concd H_2SO_4 and 112 cc H_2O , then pouring into H_2O , gave 14.1 g of crude anthraquinone 1-carboxamide-2-thioglycolic acid (VII), which was purified by recrystn from $\text{C}_6\text{H}_5\text{N}$ in which the pyridine salt is formed and from which it is regenerated by warming with 5% HCl and recrystg from PhNO_2 forming olive-green needles, m 276° . Boiled 31.5 g VII for 1.5 hrs with 11.2% NaOH , cooling, and adding HCl until acid, gave 29 g anthraquinone-1-carboxylic-2-thioglycolic acid (VIII), yellow, m $244\text{--}50^\circ$. When 42 g VIII was boiled for 15 min with 9 g NaOAc and 150 cc Ac_2O , cooled, 150 cc AcOH added, poured into H_2O , warmed, filtered and recrystd from AmOH it gave brown needles of acetyl anthraquinone-2,1-(β -hydroxythiophene), $\text{C}_{18}\text{H}_{13}\text{O}_2\text{SCH}_2\text{COAc}$, m 154° , removal of the

Ac group by boiling with 5% NaOH gave anthraquinone 2,1-(β -hydroxythiophene) (IX), $\text{C}_{18}\text{H}_{13}\text{O}_2\text{SCH}_2\text{CO}$, violet-black, m $230\text{--}40^\circ$ (decompn), when boiled with p -

$\text{BrC}_6\text{H}_4\text{COCl}$ in $\text{C}_6\text{H}_5\text{N}$ it forms the p -bromobenzoyl deriv, $\text{C}_{18}\text{H}_{13}\text{O}_2\text{SCH}_2\text{CO}_2\text{CC}_6\text{H}_4\text{Br}$,

yellow, m 228° . Anthraquinone 2,1-(α -bromo- β -hydroxythiophene) (X), obtained by adding Br_2 to IX in PhNO_2 , violet blue, m 220° (decompn). A benzylidene deriv. (XI) of IX, $\text{C}_{18}\text{H}_{13}\text{O}_2\text{SC}(\text{CHPh})\text{CO}$ prepd by condensation with BzH at

130° , golden brown, m $257\text{--}61^\circ$. Anthraquinone 2,1-thiophene-3'-indoleindigo (XII), $\text{C}_{18}\text{H}_{13}\text{O}_2\text{SC}(\text{C}_6\text{H}_4\text{ON})\text{CO}$, which was prepd in 51% yield by heating 2 g IX, 2 g

isatin α -anil and 7 cc Ac_2O for 8 hrs. at 130° , violet brown, m 290° (decompn). Similarly, 2 g IX, 1.8 g secnaphthanthraquinone, 20 cc PhNO_2 and 2 cc Ac_2O , heated for 2.5 hrs at 130° , yield 1.4 g anthraquinone-2,1-thiopheneacenaphtheneindigo (XIII), $\text{C}_{18}\text{H}_{13}\text{O}_2\text{SC}(\text{C}_{12}\text{H}_9\text{O})\text{CO}$, violet brown, m 320° (decompn). Condensation of IX and

thionaphthenequinone gave anthraquinone 2,1-thiophenethionaphtheneindigo (XIV), $\text{C}_{18}\text{H}_{13}\text{O}_2\text{SC}(\text{C}_8\text{H}_7\text{OS})\text{CO}$, brown violet, m 360° (decompn). XI, XII, XIII and XIV

form green solns in concd H_2SO_4 which is characteristic of most of the anthraquinonethioindigoids; they are not technically useful because of a lack of fastness for cotton and the poor shades of color. Oxidation of IX to bis-2,1-anthraquinonethiophenindigo, $(\text{C}_8\text{H}_7\text{O})_2$

$(\text{CO})_2\text{C}_6\text{H}_4\begin{smallmatrix} \diagup \text{CO} \\ \diagdown \text{S} \end{smallmatrix} \text{C})_2$ with persulfates in $\text{C}_6\text{H}_5\text{N}$ or $\text{K}_2\text{Fe}(\text{CN})_6$ aq. soln. was not successful; an impure product, m. above 440° , was obtained by boiling **I** in $\text{C}_6\text{H}_5\text{N}$ with piperidine or KOH (cf. Ger. pat. 461,533(1926)). **VI**. Derivatives of anthraquinone-2,1-thiazole and other products from 1-amino-2-mercaptoanthraquinone and halogen compounds. *Ibid.* 275-85.—A 63% yield of (*m*-nitrophenyl)anthraquinone 2,1-thiazole (**I**) $\text{C}_{18}\text{H}_{11}\text{O}_2\text{S}(\text{C}_6\text{H}_4\text{NO}_2)$ **N** was obtained by adding slowly with stirring 50 g. m-

$\text{O}_2\text{NC}_6\text{H}_4\text{COCl}$ to 20 g. 1,2- $\text{C}_6\text{H}_4(\text{CO})_2\text{C}_6\text{H}_4(\text{NH}_2)\text{SH}$ (**II**) in 300 g. PhNO_2 at $170-5^\circ$, continuing heating and stirring for 1 hr., cooling, filtering, washing the ppt. with hot H_2O then extg. with EtOH and recrystg. from PhNO_2 , it forms green crystals m. $319-20^\circ$ when nitrated in H_2SO_4 with fuming HNO_3 it forms dinitroanthraquinone 2,1-thiazole, m. $350-5^\circ$ (decompn.). Reduction of 2 g. **I** on boiling for 3 hrs. with 30 g. Na-S in 200 cc. H_2O and 10 cc. 30% NaOH, filtering, washing with hot H_2O and extg. with PhMe gave 1.3 g. (*m*-aminophenyl)anthraquinone 2,1-thiazole (**III**), red, m. 241° , reduction of 3 g. **I** with Na-SO₂ in d.l. EtOH and 30% NaOH gave 2.9 g. crude **III**. Red azo dyes were prepd. by diazotization of **III** and coupling with β - $\text{C}_6\text{H}_4\text{OH}$, with 2,6- $\text{C}_6\text{H}_3(\text{OH})_2\text{SO}_3\text{H}$ and with 1,3,8- $\text{C}_{10}\text{H}_6(\text{OH})_3(\text{SO}_3\text{H})$. Acetylation of **III** with Ac_2O gave (*m*-acetamidophenyl)anthraquinone 2,1-thiazole, light yellow, m. 266° . When COCl_2 was passed for 30 min. in a slow stream into 4 g. **III** in 400 cc. dry PhMe at $90-5^\circ$, cooled, filtered, washed with gasoline and dried *in vacuo* over paraffin, the crude isocyanate deriv. (**IV**), $\text{C}_{18}\text{H}_{11}\text{O}_2\text{S}(\text{C}_6\text{H}_4\text{NCO})$ **N** was obtained, m. $350-60^\circ$, the identifi-

cation as an isocyanate was established by conversion to the very insol. urea deriv.,

$[\text{C}_6\text{H}_4(\text{CO})_2\text{C}_6\text{H}_4\begin{smallmatrix} \diagup \text{N} \\ \diagdown \text{S} \end{smallmatrix} \text{C}_6\text{H}_4\text{NH}_2\text{CO}]$ (by heating **IV** with **III** in PhNO_2 for 2 hrs. at $150-60^\circ$, washing with hot AmOH and EtOH), m. above 300° , boiling **IV** with AmOH gave the amylurethan deriv., $\text{C}_{18}\text{H}_{11}\text{O}_2\text{S}(\text{C}_6\text{H}_4\text{NHCO}_2\text{Am})$ **N**, yellow, m. 221° , boiling

IV with PhNH_2 gave the phenylurea deriv., $\text{C}_{18}\text{H}_{11}\text{O}_2\text{S}(\text{C}_6\text{H}_4\text{NHCONHPh})$ **N**, which

could not be recrystd. and m. above 300° . A soln. of 2 g. **II** in 10 cc. PhNO_2 at 120° was treated with 0.3 g. cyanuric chloride in 4 cc. PhNO_2 , cooled, filtered, the ppt. washed successively with warm EtOH, H_2O , 1% NaOH, EtOH and gave 0.8 g. of the cyanuric deriv., $(\text{NH}_2)_3\text{C}_6\text{H}_3(\text{O}_2\text{S})_3\text{C}_6\text{H}_4$, red cryt. product which could not be recrystd. and m. above 210° (decompn.). Condensation of 20 g. **II** in 50 cc. H_2O , 50 cc. EtOH and 6 g. NaOH with the gradual addn. of 7 cc. $(\text{CH}_2\text{Cl})_2$ to the boiling soln., continuing the heating and stirring on the water bath for 2 hrs., filtering, washing the ppt. with warm H_2O and alc., gave 16 g. crude bis(1-amino-2-mercaptoanthraquinonyl)methylene (**V**), $(\text{NH}_2)_2\text{C}_6\text{H}_3(\text{O}_2\text{SCH})_2$, red, m. 270° (decompn.). Addn. of 5 g. **II** to a boiling soln. of 50 cc. $\text{C}_6\text{H}_5\text{Cl}$, 25 cc. H_2O , 25 cc. EtOH and 1.5 g. NaOH similarly gave 3.5 g. of a crude product from which 1.8 g. **V** was extd. with AmOH and the residue, which was recrystd. from PhNO_2 , is the β -chlorovinyl ether of **II**, $\text{NH}_2\text{C}_6\text{H}_3(\text{O}_2\text{SCH})\text{CHCl}$, red, m. 180° .

N. A. LANGE

Derivatives of 1,2-dihydroxythioxanthone. ALFRED A. LEVI AND SAMUEL SMILES. *J. Chem. Soc.* 1931, 159-8.—4,1,2- $\text{BrC}_6\text{H}_3(\text{OMe})_2$ and ClSO_3H give the 5-sulfonyl chloride, characterized as the amide, m. 226° , reduction of the chloride with Sn and HCl gives a thiol, oxidized to dinitroaryl 4-disulfide, m. 94° , with Na_2SO_3 the chloride gives the sulfonic acid, m. 122° , with concd. HI and SO_2 there results di-4-bromocetyl 5-disulfide, m. $118-9^\circ$. 2- $\text{HSC}_6\text{H}_4\text{CO}_2\text{H}$ and β - $\text{MeC}_6\text{H}_4\text{OH}$ with H_2SO_4 give 4-hydroxy-1-methylthioxanthone, yellow, m. 215° , it does not yield a diacetoborate. 1-Hydroxy-4-methylthioxanthone, deep yellow, m. 160° , results from the 1-MeO deriv. and concd. HCl, the red diacetoborate m. 235° and is decompd. by hot H_2O , this is also obtained from the 1-NH₂ deriv. through the diazo reaction and in small yields from 2- $\text{HSC}_6\text{H}_4\text{CO}_2\text{H}$ and β - $\text{MeC}_6\text{H}_4\text{CO}_2\text{H}$, 4,1,2- $\text{BrC}_6\text{H}_3(\text{OMe})_2$ and $\text{HSC}_6\text{H}_4\text{CO}_2\text{H}$ with H_2SO_4 give 40% of 4-bromo-1,2-dimethoxythioxanthone (**I**), yellow, m. 159° , HCl gives a bright red soln. 4,5-1,2- $\text{Br}(\text{HS})_2\text{C}_6\text{H}_3(\text{OMe})_2$ and 2- $\text{BrC}_6\text{H}_4\text{CO}_2\text{H}$ with K_2CO_3 in AmOH give 2-bromo-4,5-dimethoxy-2'-carboxylphenyl sulfide, m. $211-2^\circ$, concd. H_2SO_4 gives the above thioxanthone. 5,2- $\text{Br}(\text{MeO})_2\text{C}_6\text{H}_3\text{OH}$ and 2- $\text{HSC}_6\text{H}_4\text{CO}_2\text{H}$ with concd. H_2SO_4 give 25% of 4-bromo-2-hydroxy-1-methoxythioxanthone, yellow, m. 208° , it does not give a diacetoborate, concd. HCl gives the 1,2-di-HO deriv., orange, m. 212° , the Ac deriv. yields a red diacetoborate, the purple Na salt is sparingly sol. in H_2O . Hydrolysis of **I** with

coned HCl gives 4-bromo-1-hydroxy-2-methoxythioxanthone (II), orange, m 191°, also obtained from 4,2 Br(MeO)C₆H₃OH and 2-HSC₆H₄CO₂H, the red diacetoborate is hydrolyzed by hot H₂O. Cresosol and 2 HSC₆H₄CO₂H give 20% of 1-hydroxy-2-methoxy-4-methylthioxanthone (III), red, m. 173-4°, the purple diacetoborate is hydrolyzed with difficulty. The 1,2-di MeO deriv (IV), yellow, m 125°, dry HCl gives the deep red di-HCl salt. II and H₂O₂ in AcOH give the dioxide, yellow, m 243°, which gives an unstable diacetoborate. III gives a dioxide, yellow, m 190°, the red diacetoborate, m 222°, was sufficiently stable for isolation. IV gives a dioxide, yellow, m 154°. The dioxide of I, yellow, m 165°, was obtained by methylation of the K salt of the 1-HO deriv. Attempts to oxidize I with H₂O₂ gave 1,2-dimethoxythioxanthone dioxide (f), pale yellow, m 246°. The study of the basic strength of these derivs confirms the conclusions previously attained from other materials. Derivs contg the 1-HO group gave no evidence of basic character either with dry HCl or with the aq reagent. The selective attack of the 1 MeO group in demethylation is adequately explained by the structure assigned to the cation of the salts. This interpretation suggests that the 1-HO compds formed contain a similar chelate structure, which is also indicated by the insolubility of all of the 1-hydroxythioxanthones examd in aq NaOH. C. J. WEST

Derivatives of pyridine. A. BINZ AND C. RATH. IX. A new method for the preparation of 2-chloropyridines. C. RATH. *Ann.* 486, 71-80 (1931), cf C. A. 25, 953. — 2-Hydroxy-5-chloropyridine and KOH in abs EtOH, treated gradually with MeI, gave 80% of *N*-methyl-2-keto-5-chloropyridine (I), b_p 153-5°, m 44-5°, 5-Br deriv, b_p 157°, m 53° (75% yield), 3,5-di-Cl deriv, m 142° (80% yield), 3,5-di-Br deriv, m 182° (90% yield), 3,5-di-I deriv, m 227° (quant. yield). I and COCl₂ in PhMe, heated 4-5 hrs at 170°, gave 75% of 2-chloropyridine, b 163°, *N*-methyl-2-thiopyridone gives 50% of the Cl deriv. 2-Chloroquinoline, m 37-8°, is formed in 70% yields by the same method at 180°. 2-Chloro-5-nitropyridine (II), m 107°, is obtained in nearly quant. yield from *N*-methyl-2-keto-5-nitropyridine, in 90% yield from the *N*-Et deriv. and in nearly quant. yield from the *N*-benzyl deriv. 2,5-Dichloropyridine, m 60° (40% yield), 2-chloro-5-bromo deriv, m 71° (50% yield), 2-chloro-5-iodo deriv, m 99° (60% yield), 2,3,5-tri-Cl deriv, m 49-50° (90% yield), 2-chloro-3,5-dibromo deriv, m 43° (90% yield), 2-chloro-3,5-diiodo deriv, m 72-3° (50% yield). II is obtained in 25% yield by the use of SOCl₂. C. J. WEST

The preparation of 2,2'-bipyridyl by catalytic dehydrogenation of pyridine under pressure. J. P. WINAUR AND H. D. TRENK WILLINK. *Rec. trav. chim.* 50, 237-90 (1931). — In the course of expts on the action of NH₃ on pyridine at 300° under pressure and in the presence of dehydrogenation catalysts (cf C. A. 23, 5184) the formation of 2,2'-bipyridyl together with 2-aminopyridine was observed. The formation of 2,2'-bipyridyl by pyrogenic decompn of pyridine was already noticed by Meyer (C. A. 15, 3845) while Thate observed the formation of this compd. on heating pyridine under high pressure at 500°. The reaction temp. may therefore be lowered by suitable dehydrogenation catalysts and it was now found that both the Ni-Al₂O₃ catalysts described by Zelinski and Komarevski (C. A. 18, 2885) give good results. The prepn of 2,2'-bipyridyl may be carried out also by the method of Klein and Retter (C. A. 23, 143), by heating pyridine with anhyd FeCl₃ at 300°. The method described in the present paper, however, gives better results, a yield of 14-20% (relative to the pyridine which was used up) being obtained on heating 300 g pyridine with 30 g catalyst at 320-5° for 5-6 hrs. The catalyst can be used only once. W and W recommend using the catalyst described by Z and K (*loc. cit.*, p. 669 of original), which is much more easily prepd than the catalyst described on p. 667 and gives about the same results. On carrying out the same expt without the catalyst, no trace of 2,2'-bipyridyl is obtained. C. F. V. D.

The catalytic synthesis of phenylated pyridines from aldehydes and ketones with ammonia. A. E. CHENEBARIN AND D. I. OROCHKO. *J. Russ. Phys.-Chem. Soc.* 62, 1201-6 (1930), cf C. A. 18, 2495. — BzH (I) and PhCH₂CHCHO (II) react with AcH and NH₃ at 305-10° with alumina or kaolin catalyst to give a mixt. of a phenylpyridine (III) and γ-phenylpyridine (IV). I with AcMe and AcCH₃CHPh (V) with AcMe give α,α'-dimethyl-γ-phenylpyridine (VI). I and II with AcH and NH₃ in sealed tubes at 150-220° give only tar and no C₈H₈N derivs. The alumina catalyst was prepd by pptg a boiling Al₂(SO₄)₃ soln with NH₄OH, washing, drying and heating the product slightly. The kaolin should be of dense structure, low in Fe and of glistering rather than greasy fracture. The syntheses are carried out in a porcelain tube 100 cm long and 16 mm in diam, filled up to 50 cm. with catalyst. In the prepn of III and IV from I and II the following procedure was used: Add 10% HCl to the reaction mixt. consisting of an oily and an aq layer, ext. the neutral products with Et₂O, add NaNO₂ to the residue, decomp the nitroso derivs with NaOH, dissolve the free bases in Et₂O, dry, evap. the

Pt_2O and fractionate. Fraction A, b 270-80°, yielded some IV by spontaneous crystn which was recrystd twice from petr ether. The combined mother liquors were evapd to remove the solvent and again fractionally distd. The distillate was dissolved in EtOH from which III and IV were isolated as picrates, and recrystd from boiling AcMe . Fraction B, b 285-340°, yielded a small quantity of IV isolated through the picrate. The yield of III from I was 3.5%, from II, 1% of IV from I, 11%, from II, 19%. The procedure for the prepn of VI follows. The reaction mixt was acidified with 15% HCl , the neutral products were extd with Pt_2O , and the residue was cooled for 2 hrs. Sepn of crystals and a viscous oil followed. The crystals are sol in EtOH and AcMe , insol in PhH and Et_2O . Upon addn of AcOEt to the LiOH soln to slght turbidity the HCl -salt of VI was obtained and recrystd from boiling H_2O . The acid mother liquors were made alk and extd with Et_2O . After drying and evapg the Pt_2O , the free base was fractionated. The fraction b 240-95° was treated with 15% HCl in EtOH , and AcOEt added. Recovery of VI from the combined HCl salts with recrystn from petr ether gave a product m 62-63.5°. The yield of VI from I was 13%, from V, 15.5%.

LEAH V. BUTZ

The hydrogenation of pyridine with hydrogen under pressure by the Bergius process, H THAYE *Rec trav chim* 50, 77-90 (1931).—Since pyridine occurs in coal tar, while it is not present in the tar produced from coal hydrogenated with H under pressure by the Bergius process (Heyn and Dunkel, *Brennstoff Chem* 7, 208 (1926); cf *C A*, 20, 2733, 3071. Tropch and Ter Nedden, *C A* 19, 2739; Bergius, *C A* 22, 1027, 15, 3390. *Z angew Chem* 37, 400 (1921). Bube, *C A* 15, 3198) the question arose as to how pyridine behaves when submitted to this treatment. The only N compds so far hydrogenated under pressure are carbazole, which is quite stable (Piat'ev and Orlov, *C A* 23, 3697. Spilker and Zerke *C A* 21, 166), quinoline, which was converted into NH_3 and a mixt of hydrocarbons, b 60-150° (I and O, *loc cit*), aniline, which gave benzene and NH_3 (I and O, *loc cit*). The hydrogenation of pyridine with H under pressure was carried out in a rotating Fe autoclave as described by Bowen and Nash (*C A* 20, 2241), the temp being raised to 500° and the initial pressure being 50-100 atm; the heating was continued for 1-5 hrs. The gases obtained consisted of a mixt of CH_4 , C_2H_6 , C_2H_4 , C_2H_2 , C_3H_8 , C_3H_6 , C_3H_4 , C_4H_{10} , C_4H_8 , C_4H_6 (cf Tropch and Dittlieb, *C A* 19, 2793), the residue was first acidified and distd with steam and gave pentane, b 40°, BuCN , b 125-50°, a colorless liquid hydrocarbon $\text{C}_{10}\text{H}_{16}$, which however, is only present when Al or Fe chlorides are used as a catalyst, and a mixt of high boiling nitriles, b 220-60°. Finally, this soln was distd with steam after the addn of alkali and then gave AmNH_3 , unchanged pyridine homologs of piperidine, α, α' -bipyridyl and α - γ -bipyridyl. In the mouth of the manometer connection NH_4CN was deposited in the form of white platelets, further, a deposit of C was noticed. Modification of the expts gave the following results: (a) heating with H under 50 atm at 500° for 1 hr., hydrocarbons 16%, amines and nitriles 2.5%, (b) under 50 atm at 500° for 4 hrs., hydrocarbons 29%, nitriles and amines 2.5%, (c) heating under 100 atm at 500° for 1 hr., hydrocarbons 23%, nitriles 4-6%, amines very little. Pyridine was also heated at 500° under its own vapor pressure or in N or CO, with and without Al or ferric chloride as a catalyst, but no change was observed except for the formation of a trace of α, α' -bipyridyl. From these expts the conclusion was drawn that the mechanism of the hydrogenation of pyridine under high pressure is similar to the berginization of aromatic hydrocarbons and that hydrogenation precedes the cracking process. It thus follows that piperidine must be formed in the first stages of the process and as a matter of fact the same products were obtained by the berginization of piperidine. In the original paper the hydrogenation of pyridine, in general and the berginization of several classes of org compds are discussed briefly.

C F VAN DUIN

Interaction of sulfonates and sulfonamides with piperidine. FRANK BELL, *J Chem Soc* 1931, 691-15.—Turner and coworkers (*C A* 23, 2957, 24, 4006) have shown that aryl sulfonates can be very neatly severed by piperidine according to the general equation $\text{ROSO}_2\text{R}' + \text{HNC}_4\text{H}_{10} = \text{ROH} + \text{C}_4\text{H}_9\text{NSO}_2\text{R}'$. It is now found that, although this type of reaction occurs with the mononitrophenyl sulfonates, with 2,4-(O_2N) $_2$ $\text{C}_6\text{H}_3\text{SO}_2\text{C}_6\text{H}_4\text{Me}$ the nuclear link is severed, and 2,4'-dinitro-1-phenylpiperidine results. 2,4-(O_2N) $_2$ $\text{C}_6\text{H}_3\text{OH}$ gives a piperidine salt, bright yellow, m 171°, and a pyridine salt, golden yellow, m about 85°. 2,4'-Dinitro-1-phenylpyridinium *p*-toluenesulfonate is recovered unchanged after soln in AcOH but reacts vigorously with piperidine, giving 2,4-(O_2N) $_2$ $\text{C}_6\text{H}_3\text{NH}_2$. 2-(O_2N) $\text{C}_6\text{H}_4\text{SO}_2\text{C}_6\text{H}_4\text{Me}$ has no tendency to form a pyridinium compd even when boiled with $\text{C}_4\text{H}_9\text{N}$ for several hrs. Picric acid and piperidine, boiled 1 hr., give a piperidine compd., orange, m about 135°, of piperidine picrate, m 150°, the picrate was recovered unchanged after boiling with *p*- $\text{MeC}_6\text{H}_4\text{SO}_3$.

Cl for 2 hrs. A soln of mol amts of picric acid and $p\text{-MeC}_6\text{H}_4\text{SO}_2\text{Cl}$ after 12 hrs gives *picrylpyridinium p-toluenesulfonate* (I), pale yellow, m. 197° (decompn.), the Me_2CO mother liquor contained *picrylpyridinium picrate* (II). I dissolves easily in cold AcOH but 11.0 ppts. II. The warm aq soln of I deposits II on cooling. II in EtOH HCl gives *picrylpyridinium chloride* and *picryl chloride*. I undergoes vigorous reaction with piperidine but no definite product was isolated. II evolves $\text{C}_6\text{H}_5\text{N}$ with warm NaOH and when boiled with AcOH gives pyridine picrate. 4'-Nitro-*p*-toluenesulfonoxyl-biphenyl, the 3,4'-dinitro deriv., di-*p*- and *p*-toluenesulfon *p*-nitroanilide and di-*p*-toluenesulfon-*o*-nitroanilide are recovered unchanged after boiling with $\text{C}_6\text{H}_5\text{N}$. 3-bromo-5,4'-dinitro-4-hydroxybiphenyl gives only an unstable red $\text{C}_6\text{H}_5\text{N}$ salt, which decomps in the air and immediately in AcOH. 3,5,4-trinitro-4-hydroxybiphenyl gives a red $\text{C}_6\text{H}_5\text{N}$ salt, m. 198° . *p*-Toluenesulfon *p*-nitroanilide gives a bright yellow *piperidine salt*, m. 138° . The di-*p*-toluenesulfon deriv. gives a thick liquid, contg. the *p*-toluenesulfon deriv. and *p*-toluenesulfonpiperidine, m. 103° . 3- $\text{O-NC}_6\text{H}_4\text{N}(\text{SO}_2\text{C}_6\text{H}_5)\text{SO}_2\text{C}_6\text{H}_4\text{NO}_2$, on soln in piperidine, gives m. $\text{O-NC}_6\text{H}_4\text{NHSO}_2\text{C}_6\text{H}_4\text{NO}_2$ and $\text{C}_6\text{H}_5\text{NHSO}_2\text{C}_6\text{H}_5$, 2,4-($\text{O-N-C}_6\text{H}_4\text{NHSO}_2\text{C}_6\text{H}_5$) (III) and $p\text{-MeC}_6\text{H}_4\text{SO}_2\text{Cl}$ in $\text{C}_6\text{H}_5\text{N}$ give di-*p*-toluenesulfon-2,4-dinitroanilide, m. 217° . soln in piperidine gives *p*-toluenesulfonpiperidine (IV) and III. 3,5-Dinitro-4-di-*p*-toluenesulforamidobiphenyl, m. 249° , this is decomd by piperidine, giving IV. 3,5-Dibromo-4-*p*-toluenesulforamidobiphenyl, m. 196° .
C. J. WEST

The synthesis of α -hydroxy derivatives of quinoline. A. E. CHURCHMAN and A. I. KERSANOVA. *J. Russ. Phys.-Chem. Soc.* 62, 1211-16 (1930). —*o*-Toluquinoline (I), *p*-toluquinoline (II), α -naphthoquinoline (III), and β -naphthoquinoline (IV) react with dry, fused KOH at $230\text{--}60^\circ$ to give the corresponding α -OH derivs. (carbostryls) and H_2 . Isoquinoline (V) reacts at 220° . I (6.7 g) gave 3.1 g α -methylcarbostryl (VI), m. $219\text{--}20^\circ$. The reaction time was 3 hrs and increasing the time to 9 hrs. did not increase the yield. II (8 g) and 8.6 g KOH reacted 5 hrs. to give 3.3 g p -methylcarbostryl (VII), m. $232\text{--}3^\circ$. IV (13.6 g) and 10.6 g KOH reacted at 205° (9 hrs.), then at $270\text{--}80^\circ$ (3 hrs.), to give 7.3 g β -naphthylcarbostryl (VIII), slightly yellow, m. $285\text{--}6^\circ$. VIII is difficultly sol in H_2O , sol in cold AcOH, EtOH and CHCl_3 . III (8.7 g) and 10 g KOH, reacted at $270\text{--}80^\circ$ for 3.5 hrs., gave 5.1 g α -naphthylcarbostryl (IX), m. $232\text{--}3^\circ$, sol in EtOH, CHCl_3 , AcOEt, cold AcOH, insol in EtO, PhH and petroleum ether. V (2 g) and 4.7 g KOH after 4.5 hrs. gave 1.4 g α -carbostryl, m. $209\text{--}10^\circ$. Oxidation of VI and VII with alk. KMnO_4 gave the corresponding *methylaniline* and *oxalymethylaniline* acid. Similar oxidation of IX gave only a small amt. of a red ppt. which is not colored violet by alkali.
LEWIS W. BUTZ

Beckmann rearrangement with quinoline compounds. KONOMU MATSUMURA and CHUSABURO SONE. *J. Am. Chem. Soc.* 53, 1493-6 (1931). —3-Benzoyl-5-hydroxyquinoline oxime (I) and SOCl_2 in EtO give the *anilide* (II), yellowish white with silky luster, m. $211\text{--}2^\circ$ (sulfate, yellow, decomps. $211\text{--}5^\circ$), of 5-hydroxyquinoline-5-carboxylic acid (III), egg-yellow, m. 273° (Lippmann and Fleissner, *Ber.* 19, 2467 (1886), obtained this acid from 8-hydroxyquinoline, CCl_4 and KOH in EtOH) the HCl salt m. 259° (decompn.) and deposits the free base on pouring into H_2O . Ba salt, yellow needles crystg. with 5.5 mols. of H_2O , *Ac* deriv., light yellow, decomps. 312° . *Me* ester, yellow, m. $225\text{--}6^\circ$ (decompn.). I and concd. H_2SO_4 heated at 100° for 1 hr. give a sulfonated II, yellowish white, m. above 300° , gives a green color with FeCl_3 and a red color with PCl_5 . $\text{C}_6\text{H}_5\text{CO}_2\text{H}$, hydrolysis with 20% HCl gives III. The Beckmann rearrangement of 5-acetyl-5-hydroxyquinoline oxime with SOCl_2 or concd. H_2SO_4 gives 5-acetamido-5-hydroxyquinoline, m. $217\text{--}8^\circ$.
C. J. WEST

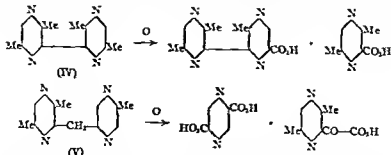
Condensation between formaldehyde and monoketones. I. Condensation of 5-acetyl-5-hydroxyquinoline with aldehydes. KONOMU MATSUMURA and CHUSABURO SONE. *J. Am. Chem. Soc.* 53, 1490-2 (1930). —5-Acetyl-5-hydroxyquinoline (I) and 40% HCHO (contg. an appreciable amt. of a ppt. of paraformaldehyde), refluxed 2 hrs., give α -methylenebis-5-acetyl-5-hydroxyquinoline (II), m. above 315° . FeCl_3 gives a green color. HCl salt, yellow needles, does not m. 300° . *Me* ester, m. $295\text{--}300^\circ$ (decompn.). Warming II, Ac_2O and AcONa on the H_2O bath 2.5 hrs. gives α -methylenebis-5-acetyl-5-hydroxyquinoline (III), m. $234\text{--}5^\circ$. I yields a *Bz* deriv., m. 168° , with HCHO this gives α -methylenebis-5-acetyl-5-benzoyloxyquinoline, m. above 315° . I and piperonal in dil. NaOH EtOH give α -piperonylidene-5-acetyl-5-hydroxyquinoline, yellow, m. $178\text{--}9^\circ$. BzMe and HCHO, with or without $\text{C}_6\text{H}_5\text{N}$, do not give a condensation product, thus the presence of an NH_2 group in the mol. may serve to facilitate the condensation.
C. J. WEST

Reduction of nitrosophenol by iron and hydrochloric acid. KONOMU MATSUMURA and CHUSABURO SONE. *J. Am. Chem. Soc.* 53, 1406-8 (1931). —5-Acetyl-5-hydroxy-

quinoline and NaNO_2 in AcOH and H_2O at 0° give the 7-NO deriv. (I), yellowish brown, decomp. $190-5^\circ$. FeCl_3 gives a light brown color, reduction of I with SnCl_2 and HCl - AcOH gives a very small yield (0.7 g. from 2.5 g. I) of the 7-NH₂ deriv., light yellow, m. $148-9^\circ$ (decompn.), picrate, orange, decomp. $201-22^\circ$; HCl salt, orange, sinters 145° . Reduction of 3.75 g. of I with Fe and dil. HCl gives 3.25 g. of the HCl salt, lemon yellow, m. $280-305^\circ$ (decompn.), of 5-acetyl-7,8-dihydroquinoline, m. $235-6^\circ$, pptn. from dil. HCl soln. gives deep red prisms, m. $241-2^\circ$ (decompn.), the colorless form crystallizes with 1.5 mols. H_2O , FeCl_3 gives a deep green color, oxime, red, decomp. $215-20^\circ$; di-*tc* deriv., m. $121-2^\circ$. 5-Nitroso-8-hydroxyquinoline (2.85 g.), reduced with Fe and HCl gives 2.85 g. of the HCl salt, orange, of 5,8-dihydroxyquinoline, m. $181-3^\circ$; sulfate, orange, decomp. 250° . Reduction of 2 g. $p\text{-ONC}_6\text{H}_4\text{OH}$ gives 1 g. $p\text{-C}_6\text{H}_4(\text{OH})_2$ and 0.02 g. $p\text{-H}_2\text{NC}_6\text{H}_4\text{OH}$. $p\text{-O}_2\text{NC}_6\text{H}_4\text{OH}$ (2.8 g.) and Fe in HCl give 0.45 g. $p\text{-C}_6\text{H}_4(\text{OH})_2$ and 0.90 g. $p\text{-H}_2\text{NC}_6\text{H}_4\text{OH}$, $o\text{-O}_2\text{NC}_6\text{H}_4\text{OH}$ (2.8 g.) gives 1.65 g. $o\text{-H}_2\text{NC}_6\text{H}_4\text{OH}$ but no other product. C. J. WEST

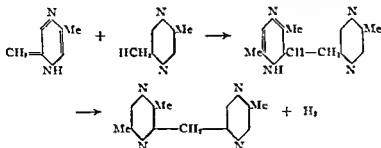
Pseudo bases. II. Equilibria and rate of change of tautomeric bases in the pyrazine series. The effect of conjugation. JOHN G. ASTON. *J. Am. Chem. Soc.* 53, 1448-70(1931), cf. *C. A.* 25, 516.—The conductivities of 1,2,5-trimethylpyrazinium hydroxide (I) and 1,2,2,5,5-pentamethyl-4-hydroxypyrazinium hydroxide (II) have been measured at 25° . Their Δ/Δ_0 values are the same at equiv. concns. and almost the same as those of LiClO_4 . The disappearance of strong base from the mixt. of I and NaOH has been found to take place by 2 reactions. One reaction takes place at a rate proportional to the square of the concn. of pyrazinium ion and the first power of the HO^- ion concn. and the other at a rate proportional to the squares of the concn. of both of these ions, when allowance is made for a salt effect according to the Brønsted theory. The same primary equil. between the quaternary base and the pseudo base hypothesized in both cases is shown to be unappreciably in favor of the pseudo base. Evidence is brought out to show that the final product is the ether of the pseudo base. The similarity between these reactions and those of pseudo bases in the pyridine and quinoline series is discussed. The equil. between II and 1,2,2,5,5-pentamethyl-6-hydroxytetrahydropyrazine has been found to be in favor of the latter. The equil. is established at a rate too rapid to measure even at 0° . The basic dissociation const. of the equil. mixt. is $K_a < 10^{-10}$. The equil. between 1,2,2,3,5,5,6-heptamethyl-4-hydroxypyrazinium hydroxide and 1,2,2,3,5,5,6-hexamethyl-6-methylenetetrahydropyrazine is established at an unmeasurable rate at 25° . The over-all base strength of the system is $K_a = 4 \times 10^{-10}$. A decompn. takes place at a rate proportional to the 1st power of the concn. of the pyrazinium and HO^- ions. The rate of formation of the pseudo base is hypothesized as the controlling step in this reaction. The soly. of 5-phenyl-10-methylacridinium hydroxide has been found to be 1×10^{-4} mol./L. at room temp. and solns. of the base appear to be highly ionized. The initial high value of the cond. on mixing 5-phenyl-10-methylacridinium chloride and NaOH is shown to be due to supersatn. The significance of the results is discussed in the light of ring conjugation, which seems to be the detg. factor in the quaternary base-pseudo base equil. C. J. WEST

The action of sodium amide on 2,5-dimethylpyrazine. A. E. CHECHERBIN and M. N. SECHUKINA. *J. Russ. Phys.-Chem. Soc.* 62, 1183-99(1930).—Amination of 2,5-dimethylpyrazine (I) proceeds very slowly with NaNH_2 . Only with strong heating in solvents is an evolution of H_2 observed, and a small quantity of 3-amino-2,5-dimethyl-



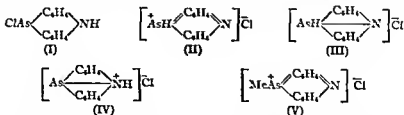
pyrazine (II) isolated. Most of the I is recovered unchanged. When the solvent is omitted, higher mol. products are obtained, the reaction takes place at lower temps.

and no H_2 is evolved. Under these conditions were isolated from the reaction mixt. some unchanged I, a small quantity of *dimethylspiperazine* (III), and 2 cryst. bases of the same empirical compn (IV), m 68° , and (V), m 135° . The structures were detd by oxidation with $KMnO_4$. In the formation of the homologous bipyrazyls, polymerization to tetramethylbipyrazyl occurs first, as is shown by the blue color of the salts formed by adding acids. As in the case of the analogous biquinolyls the color is destroyed by atm. O_2 . The formation of V has no analogy in the pyridine and quinoline series. It probably arises according to the following mechanism, involving the tautomeric form of I:




Some of the H_2 evidently goes to reduce I to III. The $NaNH_2$ either serves to produce the tautomer of I or else brings about the formation of intermediary addn. products to the pyrazine nucleus. In solvents, e. g., xylene above 145° , II was obtained in 10% yield and I was recovered to the extent of 80%. II, recrystd. from PhII, formed colorless needles, m 111° , b₁₀ 119° , sol in H_2O , EtOH, Et₂O, $CHCl_3$, PhII and PhMe. Bz deriv., m 109° . In 1 expt without solvent, 20 g I and 14.5 g $NaNH_2$ gave after Et₂O extrn. of the reaction products (4 hrs. on the H_2O bath) a red oil. This was fractionated, giving a low boiling fraction consisting of I and II, and 8 g. of a mixt. of IV and V, b₁₀ $160-75^\circ$. Further fractionation and crystn. from EtOH gave IV, sol in EtOH, and V, sol only in boiling EtOH. IV was recrystd. from petr. ether as colorless needles, m $57.5-68.5^\circ$, b₁₀ 167° , easily sol in H_2O , EtOH, Et₂O, PhII, and $CHCl_3$, difficultly sol in cold petr. ether. The aq. soln. is alk. to litmus. The soln. in strong HCl or H_2SO_4 is bright red. The $HgCl_2$ compd. was obtained as yellow prisms, m 184° (decompn.). V crystd. from EtOH in long, colorless needles, m 135° , b₁₀ $150-7^\circ$. It dissolves less easily in cold H_2O , EtOH, Et₂O, PhII and ligroin than IV, but is easily sol in $CHCl_3$. V gives no color with acids, and is neutral to litmus. The $HgCl_2$ compd. m 224° (decompn.), and the $AgNO_3$ compd. m 253° (decompn.). LEWIS W. BUTZ

The constitution of phenarsazine chloride. C. P. A. KAPPELMAYER *Rec. trav. chim.* 50, 44-50 (1931).—Previously K. showed (*C. A.* 24, 2438) that the formulation of the reaction product of $AsCl_3$ and $NHPh$, as 10-chloro-5,10-dihydrophenarsazine (I) is not in accordance with the salt-like properties and the color of this substance, the formula II, which gives an adequate account of these properties, was therefore preferred. Gibson and Johnson (*C. A.* 25, 107), however, prefer a formulation as a salt but with a bridge between the As and N atoms (III) and it is now pointed out that a decision on this point cannot be obtained from the synthesis, the sudden formation of a colored compd. at the end of the reaction pointing to an intramol. change. K. does not agree with G. and J. either on the question whether the color is to be ascribed to a tervalent or to a quinquevalent As atom, whereas G. and J. ascribe the color to the tervalent



As atom, K. holds the opinion that tervalent as well as quinquevalent As compds. may be colored or colorless, examples of both cases being known. The easy formation of colorless *N*-acyl derivs., which, according to G. and J., points to the existence or easy forma-

tion of an NH group is explained by K. by primary addn. to the N atom followed by splitting off of HCl. The *o*-quinonoid formulation is preferred by K. in view of the greater reactivity of the compd. as compared with Ph₂AsCl; moreover, in analogous compds., such as anthracene and acridine, the *o*-quinonoid formulation (von Auwers, C. A. 14, 3250, 19, 2054) has been preferred in recent years to the bridged formula. For the compd. obtained by Wieland and Rheinheimer (C. A. 15, 2037) from AsCl₃ and Ph₂NMe, the formula V is proposed. The whole subject is discussed in detail in the original paper, to which reference must be made for details. C. F. VAN DUSEN

Derivatives of the arsenic analog of 9,10-dihydroacridine. I. Wm. Gump and Hugo Stoltzenberg, *J. Am. Chem. Soc.* 53, 1428-32 (1931).—H₂N.C₆H₄.CH₂.Ph (37 g.) and As₂O₃ in the Bart reaction give 16 g. of diphenylmethane-*o*-arsonic acid, m. 161-2°. Dehydration by heating 10 g. acid in 40 cc. concd. H₂SO₄ 5 min. on the H₂O bath and pouring into 500 cc. H₂O gives quant. *acridarsinic acid*, C₁₂H₉AsO₃H. 

m. 235-6°. Reduction of 10 g. with SO₂ and HCl gives 9 g. 10-chloro-9,10-dihydroacridarsine, yellow, m. 114-5°. When dusted in the air in the smallest amts., it causes severe burning of the face, the lips and the tongue, in addn. to this, the trivalent As strongly irritates the mucous surfaces of the bronchial organs. C. J. WEST

The γ -triazines. Synthesis of two aralkylaminothio-triazines. A. Ostrogovich and V. Galea, *Ann. Acad. Lincei* 12, 162-5 (1930).—Benzylaminothio-triazine, PhCH₂C₆N₂C(NH)₂N₂C(SH)N₂, is prepd. like its isomers, the tolylaminothio-

triazines (cf. C. A. 25, 703), by condensation with heating. Recrystd. from dil. EtOH, it m. 270-1°. It is sol. in alkalis and NH₄OH, from the latter, salts of the heavy metals readily ppt., such as the Ag and Cu salts. It is also sol. in mineral acids, and gives a picrate, yellow, m. 187-8°. This compd. of all the γ -triazines prepd. gave a N₂ analysis which agreed most closely to the theoretical, i. e., 21.68% vs. 21.92% calcd. Styrylaminothio-triazine is prepd. best by condensing the reactants cold, and m. 284-5°; yield, very low. It has not been possible to prep. a Ag salt of this compd., but the Cu salt may be obtained from NH₄OH soln. by addn. of CuSO₄ and AcOH, or even by addn. of Fehling soln. The picrate m. 221-2°. A. W. COHEN

Oxonitine. THOMAS A. HENRY AND THOMAS M. SHARP, *J. Chem. Soc.* 1931, 581-2.—Oxonitine obtained by decomposition of aconitine permanganate with H₂SO₄, m. 277° (decompn.) [α]_D²⁰ -48.18° (CHCl₃, c 0.404%), gives figures on analysis agreeing best with the formula C₂₀H₂₄O₁₁N₂, the loss of the 3 C and 6 H atoms are accounted for by the loss of AcH and the Me group attached to N. C. J. WEST

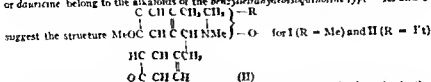
Yohimbine. H. F. MENDLIK AND J. P. WTBALT, *Rec. trav. chim.* 50, 91-111 (1931), cf. C. A. 23, 1905.—Previously M. and W. found that yohimbine can be dehydrogenated by means of Se (C. A. 23, 1905), so the present communication the compds. formed in this reaction are investigated more fully. The expts. were carried out with yohimbine, which had been purified especially by Hoffman La Roche. Hahn and Schuch (C. A. 24, 5039), however, communicate that com. preps. of yohimbine always contain isoyohimbine and, therefore, it is not certain that a chemically pure compd. has been used. An intimate mixt. of 40 g. yohimbine and 30 g. Se was heated for 30 mins. at 300°, cooled down, pulverized, mixed with an equal amt. of pure sand and first extd. with benzene for 30 min. in a Soxhlet app. and then with EtOH until the latter showed no fluorescence. The benzene ext. on cooling gave yoburine, C₂₁H₂₆N₂, m. 217°, after several crystals from EtOH, although the compd. itself is colorless, the salts and the methiodide, which show a blue fluorescence, are yellow, all these compds. are optically inactive. After sepn. of the yoburine, the benzene soln. was coned to a sirup from which dihydroyoburine, m. 170°, was obtained in colorless crystals by the addn. of hot EtOH. The alc. ext. gave the ketoyoburine, C₂₁H₂₄ON₂, in yellow crystals, m. 328° giving in EtOH and AcOH solns. characterized by a bluish green fluorescence. Probably ketoyoburine contains a CO group between 2 rings, for it does not react with PhNHNH₂ or NH₂OH. The dehydrogenation of diacetylyohimbine gives only yoburine and dihydroyoburine, not even a trace of ketoyoburine being obtained. From 40 g. yohimbine were obtained 8 g. yoburine, 7 g. dihydroyoburine and 15 g. ketoyoburine. On methylating yoburine a yellow methosulfate was obtained, which, on boiling with alkali, gave a yellow compd. C₂₀H₂₄ON₂, m. 192.5° which, from its insol. in H₂O, did not consist of the quaternary NH₂ base, but was formed most probably by simultaneous reduction and oxidation, such as is known to occur with quinaldine and acridinium bases. On melting 1 g. ketoyoburine with 5 g. KOH at 345-50°, 2 compds. are obtained,

one may be isolated by extrn of the aq soln of the reaction mixt. to which NH_4Cl has been added, with ether. It m 258° , dissolves in hot concd HCl as well as in concd KOH , but could not be investigated further, only the formula was established as $\text{C}_{10}\text{H}_{11}\text{O}_2\text{N}$. The alk. soln was then acidified and again extrd with ether after evapn of the latter a residue was obtained, which, on sublimation, gave pure *hemellitylic acid*, m. $142-4^\circ$. An authentic specimen of the latter acid was synthesized according to the method of Scholl (*Ber* 32, 3492(1900) 36, 10, 329(1903)) from *o*-xylene, Hg fulminate and AlCl_3 ; in this way a mixture of 2- and 3,4- $\text{Me}_2\text{C}_6\text{H}_3\text{CN}$ was obtained, which was saponid. to the mixt. of the corresponding acids. The latter were converted into the Ca salts, which were first sepd by the lower solv of the Ca salt of 3,4- $\text{Me}_2\text{C}_6\text{H}_3\text{CO}_2\text{H}$ and then gave a soln which, on boiling gave a ppt. of the pure Ca salt of hemellitylic acid, the solv of which decreases at higher temps. Further investigations were carried out on the dimethylindole obtained by Barger and Field (*C A* 9, 288) on heating yohimboic acid with soda lime which was not identical with any of the dimethyl- or ethylindoles then known. The odor of this indole deriv resembling that of scatole, 3 dimethylindoles were synthesized from EtCHO and the 3 tolylhydrazines according to Fischer's method. *p*-tolylhydrazine gave in this way 3,5-dimethylindole, m $74.5-5^\circ$ (picrate, m $170-80^\circ$) (cf Arbuzov and Tichwinik *C A* 4, 3218) and *o*-tolylhydrazine, 3,7-dimethylindole, b 201.2° (picrate, m $142-3^\circ$). *m*-Tolylhydrazine can give the 3,4- or the 3,6- Me_2 compd the 3,6-compd being obtained in the case of the *m*-phenetidine (Kermack, Perkin and Robinson, *C A* 16, 98), it is assumed that the dimethylindole obtained, m $116-7^\circ$ (picrate, m $179-80^\circ$) is the 3,6-deriv. None of these compds is, however, identical with the indole deriv obtained from yohimbine. The possible structures of yohimbine are fully discussed in the original paper in view of the many uncertainties and the hypothetical character of several assumptions, the original paper is to be consulted in this respect.

C. F. VAN DUSEN

Alkaloids of Sinomenium and Corcorus. XXXI. Trilobine and homotrlobine. H KONDO AND M TOMITA. *J Pharm Soc Japan* 50, 1035-59(1930), German abstr 127-9, cf *C A* 21, 2799 24, 5301-2. In the previous paper K and T proposed the formula, $\text{C}_{17}\text{H}_{19}(\text{NMe})(\text{OMe})(-\text{O}-)_2$ and $\text{C}_{17}\text{H}_{19}(\text{NMe})(\text{OMe})(-\text{O}-)_1$, for trilobine (I) and homotrlobine (II), resp. The present paper deals with the study of the properties of I and II. Oxidation of I with KMnO_4 gave a dibasic acid, $\text{C}_{17}\text{H}_{17}\text{O}_4$ (III) (*des-N-trilobinedicarboxylic acid*), m 302° , which did not depress the m. p. of the 4-methoxy-3,4'-diphenyl ether-1,1'-dicarboxylic acid, m $303-5^\circ$, of K. and Yano (*C A* 23, 4475) (III was previously described as $\text{C}_{17}\text{H}_{17}\text{O}_4$, m 288°). III on boiling with HBr in AcOH gave *des-N-desmethyltrilobinedicarboxylic acid*, $\text{C}_{16}\text{H}_{15}\text{O}_4$ (IV), m $180-1^\circ$, which did not depress the m. p. of 4-hydroxy-3,4'-diphenyl ether-1,1'-dicarboxylic acid, m $280-5^\circ$ (IV was previously reported as *des-N-desmethyltrilobinedicarboxylic acid*, m $278-9^\circ$). Methylation of IV with Me_2SO gave a compd m 302° which did not depress the m. p. of III. III when treated with CH_3N_3 gave a compd m $93-5^\circ$ which did not depress the m. p. of *di-Me-4-methoxy-3,4'-diphenyl ether-1,1'-dicarboxylic acid* (m 97°). Acetylation of IV gave a compd, $\text{C}_{17}\text{H}_{17}\text{O}_4\text{COMe}$, m 234° . Alk. fusion of III gave $p\text{-HO}_2\text{C}_6\text{H}_4\text{CO}_2\text{H}$ (m 213°) and protocatechuic acid (m 195°). *Desmethyltrilobine* (cf K. and Y., loc cit.) when treated with CH_3N_3 gave *mesomethyltrilobine* I, $\text{C}_{17}\text{H}_{19}\text{NO}_2(\text{OMe})$, m 150° , $[\alpha]_D^{25} 245.3^\circ$. V when treated with Me_2SO in alk. soln. followed by KI gave *mesomethyltrilobine-Mel*, $\text{C}_{17}\text{H}_{19}\text{NO}_2\text{MeI}$, decomp. 259° . *Mesomethyltrilobine-Me_2SO*, when treated with KOH at 100° gave *mesomethyltrilobinedimethylamine*, $\text{C}_{17}\text{H}_{19}\text{NO}_2$ (VI), m 106° , HCl salt decomp. 260° , methan. Mel (VII), m $257-8^\circ$. Oxidation of VI with KMnO_4 gave $\text{C}_{17}\text{H}_{17}\text{O}_4$, m $303-8^\circ$, which did not depress the m. p. of III. VII when treated with alc. KOH gave Me_4N , $\text{C}_{17}\text{H}_{17}\text{O}_4$, m above 300° , and $\text{C}_{17}\text{H}_{17}\text{O}_4\text{H}_2\text{O}$, m $192-5^\circ$. *Trilobine-Me_2SO*, (prepd by treating I (5 g) with Me_2SO (5 g) in CHCl_3) when treated with KOH (25 g) gave VI, m 106° . HCl salt, decomp. 259° . *meso-Mel* (VII), m $259-60^\circ$. On treating with KOH VII gave $\text{C}_{17}\text{H}_{17}\text{O}_4\text{H}_2\text{O}$ (VIII), m 192° , and Me_4N . Oxidation of VIII with KMnO_4 gave III. Methylation of VIII with Me_2SO gave $\text{C}_{17}\text{H}_{19}\text{O}_2(\text{OMe})_2$, m $90-100^\circ$. *Trilobine-Mel* (IX) when treated with KOH for two hrs. gave *trilobinedimethylamine* (X), $\text{C}_{17}\text{H}_{19}\text{NO}_2$, m 191° . On treating with MeI X gave *trilobinedimethylamine-Mel* (XI), m $206-7^\circ$. When treated with KOH XI gave $\text{C}_{17}\text{H}_{19}\text{O}_2$ (?), m 105° , Me_4N , and $\text{C}_{17}\text{H}_{19}\text{O}_2\text{H}_2\text{O}$, decomp. 214° (*Ac deriv.*, $\text{C}_{17}\text{H}_{19}\text{O}_2$, m above 300°). X or *trilobine-Mel* when reduced with Na-Hg gave *des-N-trilobine* (*mesomethyltrilobine*) (XII), $\text{C}_{17}\text{H}_{19}\text{NO}_2$, m 98° , $[\alpha]_D^{25} 195.7^\circ$. On treating with MeI XII gave VII (m 259°) which gave when reduced with Na-Hg *tetrahydro-des-N-trilobine*, $\text{C}_{17}\text{H}_{21}\text{O}_2\text{H}_2\text{O}$, m 206° , $[\alpha]_D^{25} 126.3^\circ$ (in CHCl_3); *Ac deriv.*, decomp. $144-7^\circ$, $[\alpha]_D^{25} 58.7^\circ$. Oxidation of II with KMnO_4 gave III. *Homotrlobine-Mel* when treated with alc. KOH gave *homo-*

trilobinemethylmethin $C_{15}H_{15}NO_4$, m 117°. HCl salt decomps 265°. Homotrilobinemethylmethin MeI (prepd as before) when treated with KOH gave des-N-homotrilobin (XIII) $C_{15}H_{15}O_4$, m 185-8°. Oxidation of XIII with $KMnO_4$ gave a compd (XIV) which did not depress the m p of des Δ -trilobinedicarboxylic acid (m 300°). XIV when treated with HBr in AcOH at 150° for 1 1/2 hrs gave a compd (XV), m 270-81° which did not depress the m p of desmethylides Δ -trilobinedicarboxylic acid. Methylation of XV with CH_3I gave a compd m 97° which did not depress the m p of di Me 4-methoxy 3,4-diphenyl ether 1,1 dicarboxylate. Oxidation of I, II and VI with concd HNO_3 each gave a compd $C_{15}H_{15}N_3O_4$ or $C_{15}H_{15}N_3O_5$, m 282°, contains 1 OMe and 2 C O groups, and cannot be acetylated. These facts indicate that I and II, like tetrandrine or dauricine, belong to the alkaloids of the benzyltetrahydroisoquinoline type. K. and T.



The position of R and the nature of —O— linkage are not clear. This is under further investigation. F. I. NAKAMURA

Cannabis indica resin. II. ROBERT S. CAHY. *J. Chem. Soc.* 1931, 630-8, cf. C. A. 24, 4015. — The EtO ext. of hashish of Eastern Mediterranean origin (400 g., 23%) on distn according to Wood, Spivey and Easterfield (*J. Chem. Soc.* 69, 139 (1896)) gives 90-100 g of atm. distillate, the portion non volatile in steam b₁₀ 150-270°, 147-56 g vacuum distillate, 90-130 g undistillable pitch, the 2nd fraction gives 1.4-2.0 g of hydrocarbon (I), 35-42 g, b₁₀ 100-250°, 0.5-1.10 g b₁₀ 250-85° (II) and 6-15 g residue. The fraction II simulated a pure substance, 75% of which b₁₀ 253-8° but it was impossible to eliminate the head and tail fractions. The green resin from the LiOH ext. of *Cannabis indica* of Indian origin gave similar results on distn. Portions of all fractions except that volatile in steam were separately heated at 200° for 4 1/2 hrs in an atm of CO₂ but showed no evidence of polymerization. The hashish resin yielded only about 5% to alkali and the green resin even less. II and AcCl in dry C_6H_5N give 27-8.4% of acetyl cannabinol (III), m 75°, substitution of quinoline for C_6H_5N greatly reduced the yield of III, refluxing 3 g of II with Ac₂O and AcONa for 1 hr gives only 0.1 g cryst. III after but not before crystals for seeding became available. III was unchanged by refluxing with Ac₂O. addn of fused ZnCl₂ caused decompn., bromination and nitration take place readily but cryst. derivs. have not been obtained. The only residues after sepn of the cryst. III are much more mobile than II, reacylation with AcCl and C_6H_5N gave no further yield of cryst. material. Oxidation of the higher boiling fraction with HNO_3 gave nitrocannabinolactone (6.2 g, m 176-7°, from 110 g of residues). By heating 20 g III with 6.4 g KOH in 120 cc 90% EtOH for 2 hrs. gives 90% of cannabinol (IV), b₁₀ 263-4°. Ac₂O gives quant III. Nitration of II (177 g) and of pure IV (5 g) gives the tri NO₂ deriv (V) (19.5 g and 3.3 g, resp.) which, once recrystd., never m above 150°, a purified specimen m 158°, still contained an appreciable amt. of lower nitration products or other impurities, this is similar to the product of W. S. and E., except that the Ag salt was slowly decompd by EtOH and the Na salt appeared to exist in a bright orange and a paler yellow form, the former passing into the latter in contact with H₂O slowly in the cold, rapidly on heating. Analyses of III indicate the formula $C_{15}H_{15}O_4$, Ac, of the NO₂ deriv. $C_{15}H_{15}O_4(NO_2)_3$. V as the Ag salt and MeI give 93.5% of the ether (VI), pale greenish yellow, m 150°, in MeOH only 35% of the ether is formed, 60% of V being recovered. VI is unchanged after being refluxed with EtOH-concd HCl and boiled 2 min with concd HNO_3 , it was unaffected by HI under Zeisel conditions but is hydrolyzed under the conditions for the detn of NMe. When its soln in C_6H_5N or piperidine is heated on the water bath, it rapidly assumes the red dish orange color of solns of V salts. V and $PhSO_3Cl$ in dry C_6H_5N give benzenesulfonyl trimrocannabinol m 196-7° stable to acids but very readily hydrolyzed by NaOH or C_6H_5N . The hydrocarbon I, $C_{15}H_{15}$, m 63°, it appears highly probable that the main constituent is nonacosane but some other substance is also present. C. J. W.

Barbaloin. A. GEREC. *Magyar Chem. Folyóirat* 36, 121-8, 137-42 (1930). — Léger (cf. C. A. 11, 2454, 2599) stated that barbaloin is a glucoside of aloermodin and d-arabinose in which, to account for the reducing capacity of barbaloin, the point of union could not be the 1st C atom of d-arabinose. However, the presence of a free aldehyde sugar group could not be proved by expts and the glucoside linking could not be split in any way. Léger's statement is therefore incorrect. S. S. DA FIDALY

Plant pigments. XXIX. The symmetrical lycopin formula. Perhydrolycopin. P. KARRER, A. HELFENSTEIN, B. PIEPER AND A. WETTERSTEIN. *Helv. Chim. Acta* 14, 435-8(1931); cf. *C. A.* 25, 520.—The recently proposed sym structure for lycopin has received further confirmation by O_3 treatment of the hydrocarbon, which yields 80% of the theoretical amt. of acetone, detd iodimetrically. Squalene, for which a related structure has been suggested, on similar treatment yields 90% of the acetone expected. Complete hydrogenation of lycopin should, therefore, yield the $C_{40}H_{72}$, previously prepd by a Wurtz reaction from dihydrophytyl bromide. A comparison of the substances resulting from these 2 indicated sources proved the validity of this supposition. The hydrocarbon is 2,6,10,14,18,22,26,30-octamethyltetracosane. G. ALBERT HILL.

Allomerization of chlorophyll. J. B. CONANT, S. E. KAMERLING AND C. C. STEBLE. *J. Am. Chem. Soc.* 53, 1615-6(1931).—Facts are presented which indicate that allomerization of chlorophyll is dehydrogenation brought about by the O of the air.

C. J. WEST

Ergosterol chloroacetate. MERRILL C. HART AND FREDERICK W. HEYL. *J. Am. Chem. Soc.* 53, 1413-6(1931).— α Ergosterol, prepd by the catalytic reduction of ergosterol acetate in AcOH with Pt oxide as a catalyst, contains small quantities of *allo* α -ergosterol. Ergosterol chloroacetate is very largely α -ergosterol chloroacetate with smaller quantities of the chloroacetates of β ergosterol and of *allo* α -ergosterol. Ergosterol and $ClCH_2COCl$ form a small quantity of an *addn* product, $C_{28}H_{44}O_4Cl_2$, m 129-30°, $[\alpha]_D^{25} + 8.7^\circ$. Ergosterol chloroacetate forms neither a bromide nor a peroxide nor does it react with $KMnO_4$. Upon reduction evidence of some degree of isomerization is found in the increased yield of the reduction product, *allo* α ergosterol.

C. J. WEST

Azo-chromophore (BLUMBERGER) 25. The reactions of olefins with H_2SO_4 (ORMANDY, CRAVEN) 22. Compounds of hexamethylenetetramine with certain salts of Ag and other metals and the influence of anionic volume on the capacity for association by the central positive atom (RAY, GUPTA) 6. The behavior of polypeptides built up of glycine and alanine toward polypeptidases and *N* alkali (ABERHALDEN, VON EIREN-WALL) 11A. Methylglyoxalacetic acid and its dismutation by *Bacillus coli* (VETHEL) 11A. The electrolysis of salts of *n* butyric acid (FICHTER, BÖCKMANN) 4. An x ray study of mannitol (McCREA) 2. The physical identity of enantiomers (CAMPBELL, GARROW) 2. Sulfonation (Ger. pat. 519,046) 18.

HACKEL, INGO W. D. Structure Symbols of Organic Compounds. Philadelphia P. Blakiston's Son & Co., Inc. 139 pp. \$2.50

HELLER, G. Über Isatin, Isatyd, Dioxindol und Indophenin. Stuttgart F. Enke. About 145 pp. About M. 12.70

HOUBEN, J. AND FISCHER, WALTER. Das Anthracen und die Anthrachinone mit den zugehörigen vielkernigen Systemen. Leipzig Georg Thieme. 890 pp. M. 85, bound, M. 90

Butadiene hydrocarbons. I. G. FARBENIND A-G. Fr. 696,706, June 6, 1930. Butadiene hydrocarbons are prepd by cracking higher hydrocarbons such as mineral oils, gas oils, benzene fractions and tar oils, sepg the butadienes and transforming the other products resulting from the cracking into compds such as acetone, AcH , C_4H_4 , etc., which are then used for the production of butadiene hydrocarbons. Cf. *C. A.* 25, 525

Distillation products of polymeric hydrocarbons. HENRY ARIAS. Fr. 696,812, Sept. 20, 1929. Polymeric hydrocarbons such as balata, gutta-percha, gayule, rubber, Borneo gum, etc., are distd to obtain isoprene and substitutes for linseed oil and turpentine oil.

Absorption of olefins. N-V DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ. Fr. 695,707, May 15, 1930. Olefins contg 3 or more C atoms and one or more double bonds in the mol. or their polymerization products are absorbed by strong acids in the presence of catalysts composed of one or more of the following metals, Fe, Cu, Co, Ni, Ag or metals of the Pt group or compds of any of these metals.

Absorbing olefins with sulfuric acid. H. E. BUC (to Standard Oil Development Co.) Brit. 340,098, Nov. 15, 1928. Fr. 684,567 (*C. A.* 24, 5305)

Absorbing double-bonded hydrocarbons in sulfuric acid. N-V DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ. Brit. 339,592, Aug. 6, 1929. Cyclic hydrocarbons having one double bond in a non aromatic nucleus, such as cyclohexene, are absorbed in

or combined with H_2SO_4 in the presence of catalysts such as those described in Brit 323,748 (C A 24, 1217), Brit 346,033 (C A 25, 1813), Brit 336,001 (C A 25, 1817) and Brit 337,633 (C A 25, 1813) and the products may be hydrolyzed, distilled or otherwise treated for the production of esters, alcs or ethers. An example is given of the use of cyclohexane and final production of cyclohexanol.

Recovery of liquids from gelatinized materials. I G FARBENIND A G Brit 339,610 July 10, 1929. Neutral org. liquids which have been solidified by reaction in them between NH_3 or other base and dicarbo or pumaric acid as described in Brit 340,575 (following abs.) are recovered by pressing the solid against a filter which may be assocd. with a purifier for the liquid (among which latter porous products may be used as described in Brit 340,575). Numerous details and examples are given, and among the uses to which the process may be applied is the expression of volatile combustible liquid such as pentane, ether, acetone or the like into the intake manifold of an internal combustion engine to facilitate starting of the engine.

Gelatinizing organic liquids to facilitate storage, transport, etc. I G FARBENIND A G Brit 340,115. Neutral org. liquids such as cyclohexane, benzene and its homologs, benzine, and other paraffine hydrocarbon materials, EtOAc , ether, CCl_4 , CHCl_3 , $\text{C}_2\text{H}_5\text{Cl}$ or acetone, etc. are gelatinized by forming in them a voluminous "framework" by the action of a base such as NH_3 on dicarbo or pumaric acid in the material to be gelatinized. The liquids can be recovered by pressing, by heating or by passing a current of gases over the material. Several examples are given. Cf C A 25, 2136.

Putting up organic liquids in powder form. CHEM FABRIK HENCKES A G Ger 519,418 Nov 24, 1929. Org. water sol. liquids or semi liquids are mixed with water-sol. salts which have become porous on account of loss of water of crystal. The org. liquid and the salt must be chemically inert to one another. Thus, pyridine 10 may be mixed with anhyd. Na_2CO_3 30 parts.

Nitrogenous derivatives of paraffins. I G FARBENIND A G Brit 379,962, Sept 11, 1929. Halogenated paraffins contg. more than 8 C atoms are treated with aq. or alc. NH_3 or an NH_3 -yielding agent such as $(\text{NH}_4)_2\text{CO}_3$ or urea and pressure and catalysts such as Cu or a Cu compd. may be used to assist the reaction. The halogen atoms are in part replaced by amino groups and in part split off with formation of double bonds while others remain in the mol. When alc. solns. are used, the alc. may enter into the reaction with production of ether groups. Various examples are given.

Alcoholates. ALEXANDER WACKER GRS FCS ELEKTROCHEM UND G. M. B. H. (Paul Halbig and Felix Kauffler, inventors) Ger 519,443, Jan 22, 1928. Alkali metal compds. of monohydric alcs. contg. more than 3 C atoms in the mol. are prepd. by digest. an excess of the alc. with alkali metal hydroxide until the production of alcoholate is completed. Examples are given.

Aromatic aldehydes. I G FARBENIND A G (Georg Kalischer, Heinz Scheyer and Karl Keller, inventors) Ger 519,444, Feb 26, 1927. Addn. to 514,415 (C A 25, 1536). The method of Ger 514,415 is used to introduce the aldehyde group into aromatic hydrocarbons contg. a mobile H atom. Thus, an anthracenaldehyde can be prepd. by warming C_{10}H_8 with a mixt. of formylmethylamine and POCl_3 to about 80° . Cf C A 25, 710.

Aromatic hydroxyaldehydes. J D RIEDEL-E DE HAVN A G. Ger 517,539 Feb 18, 1927. Aromatic o and p-hydroxyaldehydes are prepd. by oxidizing the alkali compds. of the corresponding propenyl derivs. in presence of excess of PhNO_2 . Thus, the m-ethyl ether of 1,2-dihydroxy-4-propenylbenzene is treated with NaOH and dried. The Na salt is then treated with excess of PhNO_2 at $120-125^\circ$. The PhNH_2 and azobenzene formed is blown out with steam and the excess PhNO_2 removed. Pptn. by HCl gives the m-ethyl ether of protocatechualdehyde. Further examples are given. Cf C A 25, 522.

Aldehyde bases. CARL MANNICH Tr 696,637, June 4, 1930. CH_3O and primary or secondary aliphatic amines are condensed with aldehydes of the general formula

$\text{R} \rangle \text{CHCHO}$ or their compds. giving rise to such aldehydes, in which R is an alkyl radical, R' an alkyl radical or H, or R and R' together represent a hydroaromatic ring. Examples are given of the prepn. of α, α -dimethyl β -dimethylamino- (b₁₁ 82*), α, α -dimethyl β -diethylamino- (b₁₁ 175-177*), α -isopropyl β -dimethylamino-propionic aldehyde (b₁₁ 66-68*) and others.

Ketones. RHEINISCHE KAMMER-FABRIK G. M. B. H. Tr 696,653, June 4, 1930. Ketones are prepd. by passing secondary alcs. in vapor form at $120-300^\circ$ and in the presence of steam over dehydrogenating catalysts. Examples are given of the prepn.

of menthone, cyclohexanone, camphor and acetophenone from menthol, cyclohexanol isobornol and methylphenylcarbinol, resp

Carbocyclic ketones. SOC. ANON. M. NAEF & C^{ie} Ger 519,446, Dec. 7, 1926 Addn to 441,273 See Brit 263,153 (C A 22, 91)

Esters. COMMERCIAL SOLVENTS CORP. Fr 696,496, May 17, 1930 Fatty acids are recovered from dil. solns from esterification reactions by extg. the acid from the aq. soln by means of the alc. which is used for the subsequent esterification

Ethanol esters. HOLZVERKOHLEUNGS INDUSTRIE A. G. Fr 696,361, May 31, 1930 Esters of EtOH are made by introducing, directly or after sepn. of impurities such as aldehydes, the mixt. of alc. and steam coming from the distn. of wort and acid (particularly AcOH) in a column system where esterification is effected. The ester and alc. in excess are sepd. and the alc. returned to the system. A suitable app. is described

Ethers of 1,2,3,4-tetrahydro-6-naphthylmethyl alcohol. I. G. FARBERND A. G. (Gustav Reddehen and Hans Lange, inventors) Ger 521,033, July 30, 1929 Addn to 516,280 (C A 25, 1841) These are prepd. by the action of aliphatic or aromatic alcs. or alcoholates on 6-chloromethyl 1,2,3,4-tetrahydronaphthalene, which is prepd. from tetrahydronaphthalene by the method of Ger 508,890 (C A 25, 716) Examples are given of the prepn. of the Me, Et and benzyl ethers b_m 133-4°, 141-3°, and 214-5°, resp.

Nitriles. I. G. FARBERND A. G. (Heinrich Hopff, inventor) Ger 517,760 Dec. 16, 1928 Benzotrichloride or its substitution products is caused to react on amide of monobasic unsatd. or satd. acids. Catalysts may be present. Thus, benzotrichloride and acetamide are heated to 150-160° to give C₆H₅CN more ZnCl₂ may be present as a catalyst. Several further examples are given

Replacing the amino group by the thiocyanato group in organic compounds. RICHARD KUHN and ERNST EICHENBERGER. Swiss 143,615, Oct. 30, 1928 Amines of aliphatic or aromatic compds. are diazotized and the N atom of the diazo group split off in the presence of a concd. soln. of thiocyanate. In the example, *o*-toluidine is diazotized and the resulting soln. treated with a satd. soln. of NaCNS or NH₄CNS to give *o*-thiocyanatoluene as a yellow oil, b_m 120°

Trithionates of secondary amines. I. G. FARBERND A. G. (Hans Lecher, Theodor Weigel and Max Wittner, inventors) Ger 517,995, Oct. 27, 1927 These are obtained by treating N-monothio derivs. of secondary amines (obtained from the amine by the action of SCl₂) with aq. H₂SO₄. Thus, N-thiodiethylamine and H₂SO₄ give the trithionate m. 143°. Further examples are given

N-Substitution products of 2,6-diaminonaphthalene. I. G. FARBERND A. G. (Richard Herz and Wilhelm Hechtenberg, inventors) Ger 517,996, Sept. 14, 1928 6-Arylamino-2-hydroxynaphthalene-3-carboxylic acid is treated with aminohydroxy compds. of the C₆H₅ series in the presence of sulfite soln. Thus, 6-phenylamino-2-hydroxynaphthalene-3-carboxylic acid, *p*-aminophenol and bisulfite soln. are heated in a reflux condenser to give 2-(4'-hydroxyphenylamino)-6-phenylaminonaphthalene, m. 205°. Further examples are given

Halosalkoxybenzanthrones. SCOTTISH DYES, LTD. Ger 516,535, Jan. 27, 1926. See Brit 256,281 (C A 21, 2989)

Decomposition of castor oil. MAURICE DEVAUX and RENÉ SORNET. Fr. 696,237, Sept. 6, 1929 Castor oil is decomposed to give a high yield of enanthaldehyde and undecylenic acid without loss of enanthaldehyde and without risk of solidification of the residual oil, by stirring the oil, heated to about 300-320°, by means of dry or superheated steam or by hot inert gases or both, which carry off the condensable gaseous products

Aromatic halogen derivatives. I. G. FARBERND A. G. Fr 695,477, May 12, 1930 Aromatic condensation products contg. a halogenated Me group besides an alkoxy or alkylthio group are prepd. by the reaction of CH₃O and hydrazides such as HCl on aromatic alkoxy or alkylthio compds., resp. A large no. of new compds. can be formed from the products as the halogen in the Me group enters easily into reaction

Vinyl halides. I. G. FARBERND A. G. Brit 339,727, Dec. 7, 1928 In effecting reaction between gaseous halogen hydrides and C₂H₂ at about 180°, a catalyst is used consisting of active C from which the activating agent has been removed, or an active C prepd. by use of H₃PO₄ and still contg. the H₃PO₄

Basic bismuth salts of organic mercury compounds. I. G. FARBERND A. G. Brit. 339,760, Nov. 15, 1929 Examples are given of the formation of reaction products from Bi nitrate and mercursalicylallylamide-O acetic acid, 4-mercurialloxy-3-acetylaminobenzene 1-arsonic acid, 3,4-mercurythiobenzimidazole-1-arsonic acid, 3,4-

mercurythiobenzimidazole 1-carboxylic acid 1,2-mercurythiobenzimidazole-4 sulfonic acid and *m*-mercuriallythiobenzobenzoic acid (and the manuf. of various of these initial lig. compds. also is described)

Metal carbamates. I G FARBEIND A G. Fr 696,374, June 2, 1930. Finely divided metallic oxides, preferably in the dry state, are treated by $\text{NH}_4\text{COONH}_2$, advantageously with heat. The production of the Zn and Ca compds. is described. Cf. C A 25, 1842

Allyl alcohol derivatives. SCHIFFING KAHNBAUM A-G. Brit. 339,882, Feb. 23, 1929. Condensation products of glycerol monohalohydrins with aliphatic ketones such as acetone, in distd. over alk. media such as caustic lime or in the presence of an org. base such as quinoline or diethylamine to split off halogen hydride and form corresponding derivs. of α -hydroxyallyl alcs. The same products also can be obtained from compds. of aliphatic ketones with other mono-esters of glycerol such as the benzene- or toluene-sulfonates. An example is given of the production of the acetone compd. of α -oxyallyl alc. from the acetone compd. of glycerol α -chlorohydrin.

Anthraquinone derivatives. IMPERIAL CHEMICAL INDUSTRIES, LTD. Fr 696,611, June 3, 1930. Halogenated phthalic acids or their salts or anhydrides are condensed with phenols in the presence of H_2SO_4 and with or without H_2O_2 . The halogen atoms in the products obtained may be hydrolyzed or replaced by OH groups. Examples given include the prepn. of 5,6-dichloro-2,6,7-trichloro- and 6,7-dichloro-quinizarin.

Benzanthrone derivatives. I G FARBEIND A G. (Willy Eichholz, inventor). Ger. 521,032 Mar. 13, 1928. Addn. to 501,083 (C A 24, 4791). Benzanthrone derivs. substituted in the 2-position with a group $-\text{CH}_2\text{R}_1$, where R_1 is H or an alkyl, aryl or other indifferent group, are prepd. by the action of alk. sapon. agents on benzanthrone derivs. substituted in the 2-position with a group $-\text{SCH}_2(\text{R}_1)\text{COR}$, where R_1 is an alkyl, aryl, aralkyl or cycloalkyl group. Thus 2-methylbenzanthrone can be prepd. by heating acetylonyl 2-benzanthronyl sulfide with KOH at 125-130°. Other examples are given also. Cf. C A 25, 2438

Isatin derivatives. I G FARBEIND A G. (Karl Thues, Ernst Runne, Karl Moldaenke and Theodor Meissner, inventors). Ger. 519,139, Jan. 31, 1926. Addn. to 451,850. Isatin derivs. trisubstituted in the benzene ring are prepd. by halogenating 6,7- or 4,7-disubstituted isatins or their bisulfite compds. The treatment of 4 (or 6)-halo-7-alkylisatins is excluded. Examples are given of the prepn. of 4,7-dimethyl-5-chloroisatin, m. 277-9°, 4-methyl-5-chloro-7-methoxyisatin, m. 251-3°, 4,5-dichloro-7-methoxyisatin, m. 274-8°, and 4,7-dimethyl-5-bromoisatin, m. 290-2°.

Lupinane derivatives. KARL WINTERFELD and PAUL HOLSCHEIDER. Ger. 521,454 Mar. 26, 1930. Bromolupinane cyanamide is prepd. by treating lupinane, $(\text{C}_{15}\text{H}_{21})_2\text{N}$, with CNBr . Dibromolupinane cyanamide is prepd. similarly from lupinine, $(\text{C}_{15}\text{H}_{21}\text{ON})_2$. Examples are given.

Naphthimidazole derivatives. I G FARBEIND A G. (Wilhelm Neelmeier and Wilhelm Meiser, inventors). Ger. 519,052, July 24, 1928. Hydroxy-1',8'-naphthoylenenaphthimidazoles and their derivs. are prepd. by fusing with alkali the 1',8'-naphthoylenenaphthimidazole sulfonic acids obtained by condensing naphthalic anhydride or its derivs. with α -naphthylenediaminesulfonic acids. Examples are given.

Pyridine derivatives. I G FARBEIND A G. Fr 696,425, April 10, 1930. *N*-Methyl compds. of the pyridine series are prepd. by the reaction of MeCl on bases of the pyridine series or its homologs, advantageously below 200° and with or without pressure and with or without catalysts, such as a chloride of Cu or Fe, and diluents such as water, ether, C_6H_6 , benzene, tetralin or decalin. Examples are given including the prepn. of the *N*-methyl chloride of α - and γ -picoline, chloropicoline and β -methoxypyridine.

Quinoline derivatives. I G FARBEIND A G. Fr 696,438, April 25, 1930. *Bz*-Nitro-Py-chloromethylquinoline and its derivs. are prepd. by treating Py-chloromethylquinoline or its derivs. with nitrating agents. Examples are given of the nitration of the 2,4-dichloro-2-chloro-4-methyl and 4-chloro-2,8-dimethyl-quinoline. Cf. C A 25, 966

Sterol derivatives. BRUNO REWALD. Ger. 520,077, Aug. 29, 1928. Sterol monophosphates are prepd. by warming ethereal solns. of sterols with P_2O_5 . Examples are given of the prepn. of the cholesteryl and sitosteryl esters.

Thioamino compounds. IMPERIAL CHEMICAL INDUSTRIES, LTD. Ger. 519,449, Mar. 30, 1930. See Brit. 331,016 (C A 24, 5765)

Condensation products from resorcinol, etc. J. D. RIEDERLE DE HAËN A-G. Ger. 521,458, Oct. 30, 1928. Resorcinol is heated with acids of the formula $(\text{CH}_3)_n(\text{COOH})_n$, where *n* is more than 2, in the presence of a condensing agent, e. g., ZnCl_2 .

Ketonic acids of the formula $(1)(3)(110)_n C_6H_4 CO(CH_2)_n COOH(1)$ are thus obtained, together with (1) benzene of the formula $O C(O)(CH_2)_n C C CH_2 CH_2 C(OH) CH_2 C -$

$O C CH_2 C(OH) CH_2 CH_2 C$, when n is 3, 5, or (2) diketones of the formula $(110)_n -$

$C_6H_4 CO(CH_2)_n CO C_6H_4 (OH)_2$, when n is more than 5. The ketonic acids can be split from the other products with Na_2CO_3 soln., and can be reduced to acids of the formula $(110)_n C_6H_4 (CH_2)_n COOH$. Homologs of these products can be obtained by starting from C alkyl resorcinols or from alkyl derivs. of the aliphatic acids. In an example, resorcinol (2 mols.) and adipic acid (1 mol.), heated to 110° with $ZnCl_2$ for 3 hrs., with stirring, yield a benzene anil $(110)_n C_6H_4 CO(CH_2)_n COOH$, m 169° , which is reducible to $(110)_n C_6H_4 (CH_2)_n COOH$, m 92° . 1 xamples of the prepn. of the following compds. are also given: $(110)_n C_6H_4 CO(CH_2)_n COOH$, m 126.7° , $(110)_n C_6H_4 (CH_2)_n COOH$, m 111° , $(110)_n C_6H_4 CO(CH_2)_n CO C_6H_4 (OH)_2$, m 186.7° , $(110)_n C_6H_4 CO(CH_2)_n COOH$, m 122° , and $(110)_n C_6H_4 CO(CH_2)_n CO C_6H_4 (OH)_2$, m 170° .

Colored condensation products. I G. FARNBERG A G. (Hans Grunowsky, inventor) Ger. 519,051, June 7, 1929. Addn. to 514,121 (*C. A.* 25, 1684). In the process of Ger. 514,121, the 1-phenyl 3-methyl-5-pyrazolone is replaced by other pyrazolones or by other heterocyclic compds. having a reactive CH_2 group, e. g., barbituric acid, rhodanic acid, sulfazone or indoxyl. Such of the products as contain $COOH$ or SO_2H groups may be used as *model dyes*; others are useful as *intermediates for dyes and drugs*. 1 xamples are given. Cf. *C. A.* 24, 6010.

Reduction products of indoxyls. I G. FARNBERG A G. (Curt Schumann, Eduard Mönch, and Bruno Christ, inventors) Ger. 516,676, Sept. 6, 1929. Addn. to 515,544. The method of 515,544 (*C. A.* 25, 2139) for producing *N*-acetyldihydroindoxyl or *N*-acetylhydriindole is modified by treating indoxyl with H in presence of catalyzers, instead of *N*-acylated indoxyl. The pH value should be between the limits 7-0.5. 1 xamples are given.

Lead salts of 2-nitroresorcinol. IMPERIAL CHEMICAL INDUSTRIES, LTD. Ger. 521,457, Feb. 2, 1929. See Brit. 312,952 (*C. A.* 24, 967) and 328,277 (*C. A.* 24, 5499).

Triphenylstibine sulfides and selenides. LUDWIG KAUFMANN. Ger. 520,306, June 11, 1927. Triphenylstibine or its homologs or derivs. are heated with S or Se , in the presence or absence of a solvent, to a temp. above the $m. p.$ of the substituted stibine. The reaction may be promoted by addn. of a catalyst such as glycerol or cholesterol, and also by actinic rays.

Thiuram monosulfides. I G. FARNBERG A G. (Ludwig Orthner, inventor) Ger. 519,445, Feb. 10, 1930. Metal salts of *N*-disubstituted dithiocarbamic acids (2 mols.) are caused to react with $COCl_2$ (1 mol.). An example is given.

Esters of phosphoric acid. I G. FARNBERG A G. Ger. 517,538, July 1, 1927. Triesters of H_3PO_4 and aliphatic or hydroaromatic alcs. are prepd. by treating the alcs. with $POCl_3$ and a solvent in the presence of sufficient amine to combine with the HCl produced. Thus, a cooled mixt. of isopropyl alc. is mixed with pyridine and CaH_2 , and added to $POCl_3$. Pyridine- HCl seps. out and can be filtered off leaving a 95% yield of triisopropyl phosphate. Further examples describe the prepn. of tributyl and tricyclohexyl phosphate.

Acid chlorides of aminonaphtholsulfonic acids. I G. FARNBERG A G. (Hugo Schweitzer and Karl Burt, inventors) Ger. 516,678, Oct. 27, 1928. Acid chlorides of 2-amino-5-hydroxynaphthalene-7-sulfonic acid and 2-amino-8-hydroxynaphthalene 6-sulfonic acid, which contain a substituent in the amino group, are prepd. by treating them with $ClSO_3H$. Thus, 2-acetylamino-5-hydroxynaphthalene-7-sulfonic acid is treated with $ClSO_3H$ with cooling to give 2-acetylamino 5-hydroxynaphthalene 7-sulfonyl chloride. Further examples are given.

Aliphatic acids. HERBERT LANGWELL. Fr. 695,724, May 10, 1930. In the manuf. of aliphatic acids and other products obtained by the fermentation of suitable substances such as cellulosic materials, the acid is neutralized as it is formed by NH_3 , which is afterwards liberated by decompn. of the salt by a non-volatile base.

Fatty acid derivatives. I G. FARNBERG A G. Brit. 310,011, Sept. 11, 1929. The polyhalogenated fatty acids prepd. as described in Brit. 330,623 (*C. A.* 25, 1842) which contain over 8 C atoms and over 2 halogen atoms in their mol. are treated with aq. or alc. solns. of the fixed caustic alkalis or with aq. solns. of weakly alk. reaction such as NH_3 , Na_2CO_3 , Na phosphate, $NaOAc$, NH_4 formate or urea. Halogen atoms are at least partially removed, forming compds. with double bonds and may be replaced

wholly or partly by OH groups. When caustic alkalis are used in the reaction, the products are sol in alkali and yield salts which are *detergents* and *emulsifying agents*. When weak alkalis are employed the products are insol in alkalis, and on boiling with NaOH polymerization occurs with possible esterification of the carboxy group, and elastic rubber like products are produced. Numerous examples are given.

Higher fatty acid derivatives containing sulfur. I G FARNHIM A G Brit 340,012 Sept 11, 1929. Solns of sulfides or polysulfides are caused to react on polyhalogenated fatty acids contg more than 8 C atoms in the mol or their esters or amides or conversion products still contg halogen obtained as described in Brit 339,675 (cf following abstr) or Brit 310,011 (preceding abstr). The halogens are in part replaced by mercapto and hydroxy groups and in part are split off with formation of double bonds. The reaction may be effected under pressure, and the products yield salts on being salted out which are typical sulfonated soaps. Various examples are given in volving the treatment of trichloropalmitic acid, hexachlorostearic acid, tetrachlorooleic acid and similar derivs. Cf C A 25, 712.

Fatty acid derivatives containing nitrogen. I G FARNHIM A G Brit 339,675, Sept 11, 1929. Halogenated fatty acids contg more than 8 C atoms in the mol and which may contain hydroxy groups and be satd or unsatd are treated with solns of NH_3 in org solvents such as MeOH, EtOH, BuOH or acetone, and Cu or a Cu compd may be used as catalyst and pressure employed to facilitate the reaction. Carboxamides are first formed and in part esters, when alcs are present. The halogen atoms are gradually replaced by amino groups and may also be split off with formation of double bonds. Examples are given of the treatment of hexachlorononoleic acid, tetrachlorononoleic acid, hexachlorostearic acid and a chlorinated oleic acid.

Reducing nitroarylsulfonic acids. CARL OCHSLEY (to Etablissements Poulenc Freres) U S 1,799,030, March 24. Aminobenzenesulfonic acids are produced by reducing a nitrobenzenesulfonic acid compd such as *p*-hydroxy-*m*-nitrobenzenesulfonic acid with Zn in an aq soln of a caustic alkali at temps of 40-70°, removing excess Zn, neutralizing the soln and sepg the pptd aminobenzenesulfonic acid.

Concentrating acetic acid. Soc DES PRODUITS CHIM DE CLAMECY and ERNEST CHARLES. Fr 696,900, Sept 19, 1929. AcOH in aq soln is concd by using 2 solvents, A and B, for the AcOH. A permits extn of AcOH from water, this solvent being slightly sol in water and dissolving only slight quantities of water. B is insol in A and can dissolve the AcOH in A and give by distn a concd AcOH. Solvent A may be C_6H_6 , xylene, CCl_4 or a mixt. of these with $CHCl_3$, PhOH, castor oil or $PhNO_2$. B may be H_2SO_4 , H_3PO_4 , a mixt of these or a concd soln of $CaCl_2$.

***p*-Sulfobenzoylbenzoic acid.** SCORRIS DIEN, LTD Ger 516,674, Dec. 1, 1926. 4'-Halogen 2 benzoylbenzoic acid or its salts is treated with a neutral or acid soln of sulfite above 100°. A small quantity of Cu salt may be added. Thus, Na 4'-chlorobenzoyl 2 benzoate is mixed with Na_2SO_3 and water in an autoclave and heated under pressure to about 180°. After some hours, the contents of the autoclave are filtered, and 4'-sulfobenzoyl 2 benzoic acid can be obtained from the filtrate. A yield of 95-100% is obtained. A further example is given.

***N*-Methyl- and *N*-ethyl-5,5-alkylphenylbarbituric acids.** I G FARNHIM A G Brit 339,037, Aug 14, 1929. Compds of this class are prepd by treating alkylphenylmalonic or -cyanoacetic esters or their derivs such as esters, amides, amide-acid esters, chlorides and nitriles, with *N*-methyl or *N*-ethyl urea or its derivs such as *N*-methyl or *N*-ethyl thioureas, guanidines and isothiourea ethers, together with other customary treatments. The same products can also be prepd by methylating or ethylating 5-phenyl or 5,5-alkylphenylbarbituric acids, or by alkylating a *N*-methyl or *N*-ethyl 5-phenylbarbituric acid, or by first alkylating a deriv of the latter compd and then converting it into the desired acid. Examples are given of the production of *N*-methyl and *N*-ethyl-5,5-phenylethylbarbituric acids, and *N*-methyl 5,5-phenylmethylbarbituric acid, and mention is also made of *N*-methyl-5,5-phenylpropylbarbituric acid.

Sulfonic acids of 1- α -aminomethylnaphthalene. I G FARNHIM A G (Wilhelm Herzberg and Hans Lange, inventors) Ger 472,923, Mar 23, 1927. See U S 1,785,935 (C A 25, 524).

Sulfur derivatives of fatty acids. I G FARNHIM A G Fr 696,500, June 2, 1930. Derivs of higher fatty acids contg S are made by the reaction of solns of sulfides or polysulfides on polyhalogen fatty acids and their derivs, if necessary with heat and pressure and with or without catalysts. Examples are given of the treatment of hexachlorononoleic acid (cf Fr 678,856, C A 24, 3517), heptachlorostearic acid, etc. The products are very reactive and are used for the prepn of new products.

Aliphatic anhydrides. HEARY DARRYS Ir 695,969, May 22, 1930. Aliphatic

acids, particularly AcOH at $350\text{--}700^\circ$, are decompd in the presence of a catalyst deposited or spread on an asbestos support. The catalysts described in Fr 613,841, 634,167 (C A 22, 3892), 667,533 (C A 24, 1124), 667,500 (C A 24, 1124) and 674,988 (C A 24, 2759) can be used Cf C A 25, 1538.

Aliphatic anhydrides. C F BOEHRINGER & SOHNE G M B H Fr 696,530, May 26, 1930. Aliphatic anhydrides and AcH are prepd by disocg with heat, in the presence of catalysts, the products of the action of the corresponding fatty acids or their salts on α,α' -dichlorodiethyl ether. Examples are given of the production of Ac_2O and propionic anhydride using ZnCl_2 or SnCl_4 as catalyst.

Acetic anhydride. C F BOEHRINGER & SOHNE G M B H Fr 696,154, May 27, 1930. Ac_2O is prepd by submitting AcOH to a pyrogenic decompn in the presence of a catalyst presenting an incomplete development of its surface, e g., a gel of SiO_2 or Al_2O_3 having an incomplete development of its surface. The porosity of the surface of the catalyst can be reduced by the action of NH_3 or HCl or by introducing C into the pores of the catalyst, or by heating it to a temp at which it begins to agglutinate.

Acetic anhydride. IMPIRIAL CHEMICAL INDUSTRIES LTD Fr 696,711, June 6, 1930. Ac_2O is prepd by a pyrogenic treatment of AcOH followed by a sepn of the Ac_2O by means of a solid absorbent such as wood charcoal.

Benzoic anhydride. I G LARSEN A G Ger 520,153 Aug 30, 1928. Two mols. of benzoyl chloride are caused to react with 1 mol. of H_2O , better yields are obtained by completing the reaction at $200\text{--}250^\circ$. A catalyst, e g., a metal or a chloride of P, may be present. Examples are given.

Acetylene. DENNIS A BRADING (one-third to Maxine Brading). U S. 1,797,400 March 24. Solid carbon is passed in finely divided state in an angular direction past an elec. arc in an atm of separately introduced H. App is described.

Ethylene. N V DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ Fr 696,561, June 2, 1930. Ethylene is absorbed in strong acids in the presence of catalysts contg one or more metals or compds of the Pt group in a finely divided state and preferably disposed on a support, or compds. of Cu, Fe, Co or Ni which if insol are solubilized by bubbling CO or NO through the liquid or if sol are converted into complex compds by means of CO or NO.

Alcohols from olefins. N V DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ Fr. 695,849, May 19, 1930. See 835,551 (C A 25, 1536).

Methanol synthesis. WM J EDMONDS and LEONARD A STENGEL (to Commercial Solvents Corp) U S 1,797,569, March 24. In effecting catalytic reaction of H with C oxides, substantially pure H is first circulated over a catalyst such as oxides of Zn and Cr and there is then added to the circulating gas a mixt. of H and C oxides in proportions required to produce MeOH and effect continuance of the reaction under suitable pressure.

Butanol and higher alcohols. N V DE BATAAFSCHE PETROLEUM MAATSCHAPPIJ Fr 695,765, May 16, 1930. See Brit. 336,811 (C A 25, 1844).

Amino alcohols. HELMUT LEGERLOIT Fr 693,675, Mar 15, 1930. Aromatic amino alcs. and their derivs. of the type $\text{XOC}_6\text{H}_4\text{CH}(\text{OH})\text{CH}_2\text{NRR}'$, in which X, Y and R denote an atom of H, a univalent hydrocarbon group or an acyl group and R' a univalent hydrocarbon group, are prepd. by the reaction of H under ordinary or increased pressure on the corresponding ketonic compds. as a soln. of the base or its salts in an appropriate solvent such as water, EtOH or MeOH and in the presence of a catalyst such as metals of the Pt group, Ni, etc.

Ethylene glycol. ALBERT MAIER Fr 697,171, June 10, 1930. Polyales., particularly ethylene glycol, are prepd by saponig org esters, preferably acetates, isolated after prepn. in known manner from appropriate halogenated hydrocarbons, by water in excess, under pressure, at temps. above the b p.

Acetaldehyde from acetylene. CONSORTIUM FÜR ELEKTROCHEMISCHE INDUSTRIE G M B H (Erich Baum and Martin Mugdan, inventors) Ger. 517,893, July 16, 1914. CH_3CHO is formed by the action of a hot soln. of 6 to 35% H_2SO_4 on C_2H_2 in the presence of Hg compds as catalysts. The C_2H_2 is circulated in excess through the app., the proportion of absorbed C_2H_2 to unabsorbed C_2H_2 being kept so that no material alteration of temp occurs during the reaction. Agents to absorb or polymerize the formed CH_3CHO are present. An example is given Cf C A 25, 2441.

Monoethanolamine dinitrate. DYWIDET A G FORM ALFRED NOBIL & Co (Phokion Nauom and R. von Sommerfeld inventors) Ger 514,955, Dec. 6, 1929. A soln. of monoethanolamine or monoethanolamine mononitrate in concd HNO_3 is introduced into alc. with cooling. Examples are given stating that 76–90% yields are obtained.

Monoethanolamine dinitrate. DANSMIT A G NORM ALFRED NOREL & Co (Phokan Naum and Robert von Sommerfeld, inventors). Ger 517,832, Apr 12, 1930. Addn to 514,935 (preceding abstr.) $\text{NH}_4\text{C}_2\text{H}_4\text{OH}$ is dissolved in excess of concd HNO_3 and the excess acid distd off *in vacuo*. The residue is added to alc. or alc-ether mixt with cooling, to ppt the dinitrate. Examples are given in which 88 and 97% yields are claimed.

Hexamethylenetetramine. S KARTEN & BROS. Ger 521,476, April 21, 1927. A mixt of $(\text{CH}_3)_3\text{N}$, and NH_4Cl , such as is obtained by reacting CH_3Cl with NH_3 , is sepd into its components by extrn with a cold soln of NH_4Cl . $(\text{CH}_3)_3\text{N}$, is pptd from the ext. by means of NH_3 , and washed with a concd soln of NH_3 . The extrn with NH_4Cl soln is effected in a series of containers, the ext. from one batch of mixt. being passed to the container for the next batch.

Hexamethylenetetramine. S KARTEN & BROS. Ger 521,475, Oct 29, 1925. See Brit 240,416 (C A 21, 414).

Oxidizing tetrahydronaphthalene. J D RIEDER-F DE HAEN A G (Ludwig Hess and Simon Felsler, inventors). Ger 520,291, Apr 24, 1925. α -Keto-tetrahydronaphthalene and other oxidation products of tetrahydronaphthalene are prepd by treating the latter at a temp. below 130° with O or gases contg O, preferably in the presence of a catalyst, e. g., a salt or oxide of Mn or Cu. Examples are given.

1-Alkoxybenzene-3-thioglytol-4-carboxylic amide. I G FARRINGTON A G. (Norbert Steiger, Erwin Hoffa and Hans Heyna, inventors). Ger 516,994, Apr 13, 1927. Addn. to 514,505 (C A 25, 2156). The method of 514,505, is used for the prepn of the above, using 1-alkoxybenzene-4-cyano-3-sulfonyl chloride as the starting material. This is treated with a metallic reducing agent in the presence of a strong mineral acid and an indifferent org. solvent, and condensing the resulting 1-alkoxybenzene-4-carboxylic amide-3 mercaptan with ClCH_2COOH . In an example, 1-methoxybenzene-4-cyanobenzene-3-sulfonyl chloride is reduced by Zn dust and HCl to give the corresponding mercaptan which is then treated with ClCH_2COOH to give 1-methoxybenzene-3 thioglycol-4-carboxylic amide, m. 208° . A further example is given.

Oils containing cineole. RHEINISCHE KAMPFER-FABRIK G M B H. Fr 696,424, Apr 9, 1930. Ger 519,447, Apr 16, 1929. Mixts of oils contg cineole are prepd by treating terpinol hydrate, terpinol, terpineol or like compds which have the same empirical compn as cineole, or a higher water content, at temps. below 80° with dehydrating agents such as H_2SO_4 , H_3PO_4 , sulfonic acids, etc. The mixt. of oils can be extd by indifferent agents such as CHCl_3 , PhMe , etc.

Camphor. I G FARRINGTON A-G (Otto Schmidt, inventor). Ger 520,224, June 25, 1927. Camphor is prepd by treating borneols at a raised temp. with mixts of dithydrogenating catalysts with metal oxides other than alkali or alk. earth oxides. Thus, vaporized borneol and isoborneol can be passed at 350° over a contact mass prepd by heating a mixt. of NiCO and MnO to H_2 .

Menthol. HANS JORDAN (to Schering Kahlbaum A-G). U. S. 2,797,612, March 24. See Ger 512,719 (C A 25, 1260) and 514,594 (C A 25, 2157).

Inactive menthol. RHEINISCHE KAMPFER-FABRIK G M B H (Karl Schillkopf, inventor). Ger 516,651, Oct 11, 1923. Addn to 499,819 (C A 24, 2145). The racemic liquid mixt. of isomeric menthol is sepd. by direct treatment with H under pressure in the presence of catalysts. In the example, the liquid mixt. is heated to 200° with H under a pressure of 5-30 atm. in an autoclave, in the presence of a Ni catalyst. The catalyst is removed by distn. or filtration and the inactive menthol (of higher m. p.) is sepd. by freezing and centrifuging. Mixts of thymol isomers can be sepd. by the same method.

Acrolein. SCHERING-KAHLBAUM A G. Fr 695,931, May 21, 1930. Acrolein is made from glycerol using as catalysts salts of acids having at least 3 acid functions or mixts. of these salts. Examples are given of the use of phosphates of Fe and La.

Polyglucosans. KARL FREDENRACEN and BURCKHARDT HELPERICH. Ger 521,340, July 14, 1929. Mono- or di-saccharides are treated with practically anhyd. HI . Examples are given.

Morpholines. IMPERIAL CHEMICAL INDUSTRIES LTD. Ger 520,156, July 6, 1928. See Brit 298,336 (C A 23, 2723).

Styrenes and polymers. THE NAUGATUCK CHEMICAL CO. Fr 695,575, May 14, 1930. A halogenated alkylbenzene and an org. base such as pyridine are heated together under a pressure not above normal. Examples are given. Fr 695,576 describes the prepn. of compact non friable polymers of styrenes by heating the styrene at atm. pressure in soln., in the presence of a substance chosen from among a group contg alkali hydroxides, alkali salts, org. bases and water. Cf C A 25, 1539.

11—BIOLOGICAL CHEMISTRY

PAUL E. HOWE

A—GENERAL

FRANK P. LINDERHILL

The solubility of phosphatides. BRUNO REWALD *Algem. Ch. Fernz.* 27, 363-4(1930).—After fresh egg yolk was extd. with acetone, it was found that the ext. was richer in lecithin than the residue. The solv. of the lecithin is probably due to the complex mixt. of acetone, egg oil, lecithin and water. For a true estn. of lecithin, the acetone ext. must be analyzed for P as well as the alc. ext. of the residue.

W. F. BOLLENS

Temperature and life and death. A. T. CAMERON *Trans. Roy. Soc. Can.* 24, 302-3, 53-93, 1930.—A review. A. T. C.

Action of proteolytic enzymes on crystalline insulin. A. F. CHARLES AND D. A. SCOTT *Trans. Roy. Soc. Can.* 24, Sect. V, 95-9(1930).—Cryst. insulin cannot be broken down into a simpler physiol. active substance by the proteolytic action of pepsin or trypsin. A. T. CAMERON

Studies in cholam diastase. I. Electrodialysis and electroosmosis of cholam diastase. D. NARAYANAMURTI AND ROLAND V. NORRIS. *Proc. 15th Indian Sci. Congr.* 1923, 166.—By electrodialysis cholam malt diastase can be purified and considerably increased in activity. The ash and protein contents are much reduced by this method. An attempt was made to test the two-enzyme theory of diastase by electroosmotic expts. Preliminary expts. conducted in a 5-celled app. indicate some evidence in favor of the theory, the ratio, liquefaction power and saccharification power being different in the different fractions. E. J. C.

Kinetics of diastatic action. D. NARAYANAMURTI AND ROLAND V. NORRIS. *Proc. 15th Indian Sci. Congr.* 1923, 166.—In expts. on the hydrolysis of potato, rice and cholam starches and ovster glycogen by cholam diastase, glycogen, in agreement with other observers, is found to be the most resistant. E. J. C.

The behavior of polypeptides built up of glycine and alanine toward polypeptidases and normal alkali. EMIL ABDEKHALDEN AND ELLAVON FRIEDRICH. *Fermentforschung* 12, 378-410(1931).—Evidence thus far obtained seems to indicate that the simple amino acids, glycine and alanine, do not occur successively but alternately in the peptide chains of proteins. Polypeptides built up exclusively of glycine are more or less resistant to trypsin kinase and erepsin, depending on the length of the chain. Di- and tripeptides of alanine are amenable to erepsin but not to trypsin kinase, and longer chains have not yet been tested. The present paper deals with polypeptides in which there is an alternation of glycine and alanine. The usual method of synthesis was employed, and a no. of the peptides were converted into the PhNCO, β -C₁₂H₂₁SO₃ and Bz derivs. The products were subjected to the hydrolytic action of erepsin, trypsin-kinase and N NaOH. Chloroacetyl-d-alanine (I) \rightarrow glycyl-d-alanine (II), $[\alpha]_D^{25} = -42^\circ$ (NH₄Cl double salt, $[\alpha]_D^{25} = -9.4^\circ$) \rightarrow d- α -bromopropionylglycyl-d-alanine, decomp. 192° , \rightarrow d-alanylglycyl-d-alanine (III), decomp. above 259° , \rightarrow chloroacetyl-d-alanylglycyl-d-alanine (IV), sinters 191° , \rightarrow glycyl-d-alanylglycyl-d-alanine (V), browns 231° , dl-Alanylglycine (VI) \rightarrow chloroacetyl-dl-alanylglycine (VII), decomp. $180-7^\circ$, \rightarrow glycyl-dl-alanylglycine (VIII), m. 243° , \rightarrow dl- α -bromopropionylglycyl-dl-alanylglycine (IX), decomp. 183° , \rightarrow dl-alanylglycyl-dl-alanylglycine (X), decomp. 200° , II \rightarrow chloroacetyl-glycyl-d-alanine (XI), m. 157° , \rightarrow glycylglycyl-d-alanine (XII), browns 215° , \rightarrow dl- α -bromopropionylglycylglycyl-d-alanine (XIII), m. $160-9^\circ$, \rightarrow dl-alanylglycylglycyl-d-alanine (XIV), m. 203° . Similar syntheses through the amino acid chlorides also were undertaken, but with poorer yields. I + PCl₅ \rightarrow chloroacetyl-d-alanyl chloride (XV) + II \rightarrow V. NH₄ salt of I in Et₂O + POCl₃ \rightarrow XV + II \rightarrow V. II in AcCl + PCl₅ \rightarrow XV + II \rightarrow V. d-Alanine \rightarrow d-alanyl chloride + II \rightarrow III \rightarrow IV, $[\alpha]_D^{25} = -119^\circ$, \rightarrow V. The Bz deriv. of V m. 174° , the PhNCO deriv. of: II m. 151° , III m. 148° , V m. 178° (decompn.), VI m. 203° , VIII decomposes 208° , X m. 144° , XII m. 176° (decompn.), XIV m. 162° , the β -C₁₂H₂₁SO₃ deriv. of III m. 182° (decompn.), V decomps. 215° , VI m. 175° , VIII m. 213° (decompn.), X sinters 202° , XII sinters 222° , browns 245° , XIV decomposes above 257° , the dr- β -C₁₂H₂₁SO₃ deriv. of II decomposes 250° . Cu salts were prepd. from II, III and V and their optical rotations detd. with a quartz-Hg lamp. The 2 dipeptides II and VI were hydrolyzed at practically the same rate by N NaOH at 37° , and the tripeptides III, VIII and XII at different rates; the last was hydrolyzed the most

centrifuge it is necessary to know the partial sp. vol. A graphic method is described which is an extension of the method of Lewis and Randall for the calcul. of partial molar vol. from pycnometric data. N. A. LANGE

Studies in iodized casein. I. Preparation of iodized casein. SINGHVI MASTRI *Acta Sci. Ind. Imp. Nep.* 13, 24-30 (1931)—To obtain iodized casein with the max. content of I M. treated an alk. soln. of casein with a soln. of I_2 -KI in NaOH, Na_2CO_3 , $KHCO_3$ and $NaHCO_3$ solns. at 0, 15, 35-43, 60-65 and 80-85°. The iodized casein was dissolved in an excess of NaOH and pptd. with $NaHSO_4$ and H_2SO_4 . It was purified by pptn. and dried with the aid of alc. and ether. The final product was tested for protein reaction and I content. The KI and Na_2CO_3 method at 15-40° proved most appropriate for M's purpose. II. The behavior of iodized casein toward proteolytic enzymes. *Ibid.* 271-83—The iodized casein was hydrolyzed by pepsin, trypsin and tissue proteases. afterward half of its I could be pptd. by $AgNO_3$. III. The behavior of iodized casein in the dog. *Ibid.* 284-6—When the iodized casein was fed to dogs (29-44% of the I was recovered in the urine in the form of many iodides. Org. compds. contg. I if present in the urine, were so scarce that they could not be isolated. B. S. LEVINE

The lipoids with a special review of the newer developments in lipid research. H. W. M. JONES. *Exptl. Physiol.* 31, 165-235 (1931)—A review. C. M. M.

Potential differences across natural membranes separating unlike salt solutions. S. C. BROOKS, A. C. GREST AND R. J. GREST. *J. Exptl. Biol.* 2, 124-32 (1931)—When the lower epidermis of the bull's scale of the onion is used as a membrane to sep. 0.1M and 0.01M solns. of KCl, NaCl or LiCl a transep. p. d. occurs. The magnitude decreases in the order $K > Na > Li$. The p. d. of 0.1M 0.01M $CaCl_2$ is small and opposite in sign. If the epidermis sep. like concns. of K, Na and Li a low potential arises. When K salts with different anions are used in the cells, little p. d. is observed. The epidermis seems to be a measure of anion and cation permeable areas. C. M. M.

The arginase law and the formation of urea in the autolysis of the liver of vertebrates. A. CARMIGNI. *Bull. ex. sci. nat. ser. 5*, 1142-4 (1931)—The aim was to see whether a close parallelism existed between the presence of arginase in the liver and the formation of urea during the autolysis of the liver of various classes of vertebrates. Only in the liver of those vertebrates in which arginase is absent (birds, Ophidiidae) was there a const. and definite lack of urea, whereas in the liver of those vertebrates in which arginase is present (mammals, Cheloniidae, Amphibians, fishes), urea was formed during the autolysis. These results indicate the absence of the "arginase-arginase" system, the formation of urea in the liver does not take place, a fact, which supports the hypothesis that the formation of urea in the autolysis of liver is basically bound to the "arginase-arginase" system and not to other factors. PETER MARCOTTA

Biochemical investigations on the cerebrospinal lipoids. GISELA BLANKALANT AND RENATA SOCIAL. *Arch. Biol.* 28, 267-6 (1933) *Physiol. Abstracts* 15, 633, cf. C. A. 24, 4571—The influence of species, age and alimentation on the cerebrospinal lipoids has been studied. A. D. MEYER

The optical alterability of bilirubin solutions. EMILIO BECCARI. *Arch. Biol.* 28, 452-60 (1933), cf. C. A. 24, 3235—The changes (fading) are due to O₂, light, pH and enzymes. It is never observed in pure solns. A. E. MEYER

The catalase in human blood in its relation to the season. MARIO RIGNONI. *Arch. Biol.* 28, 482-9 (1933)—The catalase content in the blood has a max. in spring and in fall. The temp. is not the only influencing factor. A. E. MEYER

The presence of urease in the human gastric mucosa. MARIO RIGNONI. *Arch. Biol.* 28, 474-8 (1933), cf. C. A. 24, 2787—Urease is found in the human gastric mucosa where it is supposed its function is to liberate NH_3 which is changed to NH_4Cl in the blood and serves as the material out of which HCl is then produced by the gastric glands. S. MOROGLIS

Biological oxidation-reduction. H. VON EYLER AND R. NILSSON. *Stand. Arch. Biol.* 59, 201-16 (1931)—The transformation of polysaccharide and hexose to hexose diphosphate ester in buffered mixts. is discussed and the role played in this process by the cozymase is pointed out, as well as the effect of adding AclI to the system. The possible action of cozymase in the synthesis of higher carbohydrates is considered. The carbohydrate stored by yeast living on galactose has been studied specially and it was found that it probably yields Robison's ester on hydrolysis. S. MOROGLIS

Reciprocal relationship between cholesterol and some protein fractions. V. N. NIKOLAYEV. *Biochem. Z.* 232, 50-7 (1931)—A certain portion of the cholesterol of normal horse and dog plasma is firmly bound to the globulin fraction and to a part of the fibrin. A similar combination has been found in certain fractions of hen egg pro-

tein and in a part of the horse serum fractions precipitable by tungstic acid after preliminary removal of the globulin and albumin fractions. Cholesterol forms a less firm combination with the albumins, a portion of the fibrin, of the egg protein and of the same horse serum fractions. S. MORGULIS

Behavior of phosphate buffer mixtures with different cations. S. M. NEUSCHLOSS AND R. P. RIZ INÁREZ *Biochem Z* 232, 106-22(1931)—It is shown that the classical formula does not correctly express the H-ion concn of phosphate mixtures, and a different formula is suggested which reproduces more closely the exptl values. In mixts of $\text{NH}_4\text{H}_2\text{PO}_4$ and $(\text{NH}_4)_2\text{HPO}_4$ the p_{H} is always less than in corresponding mixts of the K salts. This difference is due, for one thing, to the diminished activity of the HPO_4^{--} ions in the presence of NH_4^+ ions and secondly, to the hydrolysis which $\text{NH}_4\text{H}_2\text{PO}_4$ undergoes in relatively alk solns. A buffer equation is proposed which gives satisfactory results with mixts of the NH_4 salts. The difference between the p_{H} of K or NH_4 phosphate mixts increases with rising concn of the phosphate or with the addn of NaCl to the mixts. The osmotic pressure and conductivity factors of $(\text{NH}_4)_2\text{HPO}_4$ are less than those of K_2HPO_4 of the same concn. S. MORGULIS

Further studies on amylotransferases. S. NISHIMURA *Biochem J* 232, 156-64 (1931), cf *C A* 25, 2449—The higher dextrans can be synthesized enzymically, the synthetic product being more easily hydrolyzed by malt amylase than the simple dextrans from which the higher form was synthesized. The more highly the dextrans are polymerized, the closer does their hydrolysis resemble that of starch. S. M.

The absorption spectrum of bilirubin in chloroform, alcohol and alkalis. LUDWIG HILLMAYER *Biochem Z* 232, 229-39(1931)—Bilirubin in soln in CHCl_3 or in EtOH shows a very definite absorption band with a max at 450μ , diminishing continually from this point even into the ultra violet range of the spectrum. In alkali bilirubin manifests continuous absorption in the visible range, but the absorption curve cannot actually be deid on account of the extreme instability of the alk soln. This is due not to an oxidative alteration of the bilirubin but simply to the alk solvent. S. M.

Glucoseoxidase. IV. Glucoseoxidase from *Aspergillus niger*. Behavior toward disaccharides (maltoseoxidase), glucuronic acid and ethyl alcohol, experiments with methylene blue and moniodoacetic acid. D. MÖLLER *Biochem Z* 232, 123-31 (1931), cf *C A* 24, 2769—Preps of glucoseoxidase do not act on lactose, but they do affect sucrose, the glucose formed from its hydrolysis being oxidized to glucuronic acid. When the enzy prep is heated for 30 min at 70° the sucrase is destroyed but not the glucose oxidase, and it no longer acts upon the sucrose. The effect upon maltose is direct, the maltose actually being oxidized by the prep, but this is due not to the glucose oxidase but to a different enzyme, maltoseoxidase, which is more thermostable than the former and oxidizes maltose with atmospheric O_2 to a non reducing substance. Glucoseoxidase does not act either upon glucuronic acid or EtOH , its oxidative function being apparently limited to compds with the same configuration on the C atoms 1 and 6 as in glucose. $\text{CH}_3\text{CO}_2\text{H}$ injures the activity of the glucoseoxidase much less than that of zymase. Preps of glucoseoxidase contain only maleic acid isohydrazide, as they promote the reduction of methylene blue only by maleic acid. This reductase action is inhibited by glucose, fructose, glucuronic acid, etc. S. M.

Methylglyoxalylacetic acid and its dismutation by *Bacillus coli*. SIG VPIKRI. *Biochem Z* 232, 435-41(1931)—Evidence is brought forward to show that *B. coli* accomplishes almost quantitatively the dismutation of methylglyoxalylacetic acid to α -hydroxyglutaric acid. This is an asym biosynthesis, since exclusively the dextrorotatory acid, in a practically pure condition, appears. S. MORGULIS

Comment on Lustig's paper—"Studies on the concentration of pepsin and chemistry of its action." I. A. SHORODINTSEV AND A. N. AONVA *Biochem Z* 224, 471-3 (1930)—It is shown that Lustig's results (*C A* 24, 636) do not contradict, but on the contrary corroborate, the previous findings of S. and A. S. MORGULIS

The order of death of organisms larger than bacteria. OTTO RAHN, *J. Gen Physiol* 14, 315-37(1931), cf *C A* 24, 2489, 3530—In a previous publication, R. showed that the logarithmic order of death of bacteria can be accounted for by assuming the presence of some very unstable mols so essential for reproduction that the inactivation of a single one of these mols prevents cell reproduction and makes the cell appear dead according to the standard method of counting bacteria. In the present paper the motile alga, *Chlamydomonas*, is the only one of the larger organisms dealt with which has a logarithmic order of death. Probably more than 1 mol must be destroyed to kill a yeast cell. The following cells or organisms give curves (log of survivors plotted against time) which suggest that the no. of essential reacting mols is

$\frac{1}{2}$ N alkali. The diazo reagent is about $\frac{1}{2}$ N acid because of HCl. On mixing equal vols of the 2 a weakly acid soln is obtained buffered by NaH_2PO_4 . The method was tested with a standard bilirubin soln in chloroform with 1.0 to 0.1 cc of the bilirubin soln. Ninety six % alc. was added to make each vol 10 cc and to this 2.5 cc of the phosphate buffered diazo reagent was added. The color produced from each was read in a Duboscq colorimeter against a standard contg 0.05 mg bilirubin. A graph from the results shows a strict linear proportionality between the amt of bilirubin present and the amt of color produced from it by the diazo reagent. For routine work in examg serum for bilirubin to 1 cc of serum is added 0.5 cc of diazo reagent and 2.8 cc of 96% alc., giving a diln of approx 1 in 4. It is preferable to express the concn of bilirubin in the serum as mg per 100 cc.

HARRIET F. HOLMES

A method for the cataphoresis of biological materials. F. C. SMITH AND J. MAR-
RACK *Brit J Exptl Path* 11, 492-4(1930).—An app for the cataphoresis of biol materials is described. It was designed to afford (1) a continuous path of buffer soln for the current passing between the electrodes, so that no agar bridges need be employed, (2) uniformity of the bore of the tube connecting the vessel contg the electrodes, so that the potential gradient across the moving boundary can be easily measured, (3) ease of filling and forming a sharp boundary with small vols (2 or 3 cc.) of the suspensions to be tested, (4) ability to withdraw the whole sample should it be needed for other investigations, (5) a method of obtaining small samples during cataphoresis without disturbing the boundary, so that the effect of the reversal of polarity could be studied when the boundary is invisible (e.g., virus suspensions), (6) ease of washing out with cleaning fluids.

HARRIET F. HOLMES

Studies on the biochemistry of sulfur. IV. Colorimetric estimation of cystine in casein by means of the 8-naphthoquinone reaction. M. X. SULLIVAN *U S Pub Health Repts Suppl* No 78, 13 pp (1929), cf *C. A.* 20, 2185, 23, 3941, 4445.—*Procedure for the hydrolyzate*—Various amts of casein from 1 to 10 g have been hydrolyzed. The 5 g samples are best, in general, since the final vol of soln prepd for colorimetric work is sufficient to give several tests and to allow comparison with the Okuda (cf *C. A.* 18, 2014, 3613, 3400, 20, 1034, 1252) and the Folin-Looney methods (cf *C. A.* 16, 1790). The 5 g of casein and 15 to 20 cc. of 20% HCl are put into a small acetylation flask, fitted with a ground glass reflux. The flask is then placed in a crisco bath, the heating started and hydrolysis run for 6 hrs, the count begins at the time the bath reaches 125°. The inside temp of the hydrolyzate runs about 100°. The black mixt is poured out into a 100 cc. beaker, and, by washing with distd water, made to about a 10% acid. The combined mixt. is decolorized by warming with purified (acid extd.) decolorizing C, 2 g in case nitrite is used, as was S's practice, or, better, 0.5 g carbottanin, which is now his practice. The mixt. is filtered on a small Buchner funnel, and the residue on the paper is extd with 15 cc of hot N HCl and then washed with 15 cc. of cold N acid. The combined filtrates are neutralized with 5 N NaOH added dropwise with stirring until the pH is about 3.5 (that is, yellow to thymol blue) and made to 100 cc by means of 0.1 N HCl. With small amts of material, only 1 g of casein has been used. The sample is mixed with 5 cc of 20% HCl and put into the oil bath and hydrolyzed 6 hrs. The mixt is poured out and, by washing with distd water, made to 10 cc, decolorized, neutralized and brought to about 25 cc with 0.1 N HCl, and the colorimetric detn made as soon as possible. Later, work with 1 g samples, with and without the formation of humin, will be reported. *Colorimetric estn*—To 5 cc. of the neutralized hydrolyzate, brought to a definite vol and with the temp of the soln not under 20°, add (A) 2 cc of 5% aq NaCN, mix, and wait 10 min; then add (B) 1 cc. of 0.5% aq sodium 1,2-naphthoquinone-4-sulfonate and mix by shaking for 10 min; next add (C) 5 cc. 10% Na_2SO_3 in 0.5 N NaOH, mix, let stand 30 min, add (D) 2 cc. 5 N NaOH, mix and finally add (E) 1 cc. of a 2% $\text{Na}_2\text{S}_2\text{O}_8$ in 0.5 N NaOH. Compare with 5 cc of a cystine standard, 100 to 200 p.p.m. in 0.1 N HCl, similarly treated. The standard and unknown should be close together in colorimetric reading. On adding the $\text{Na}_2\text{S}_2\text{O}_8$ the brown red color developed in the cystine-contg soln is changed to a more vivid red, if cystine is present to the extent of 100 p.p.m. or over. Weaker solns tend toward orange. Other substances in protein hydrolyzates may show color in the first stages of the reaction, but this color in the absence of cystine or cysteine is discharged to yellow by the hyposulfite. When the standard was pure cystine directly dissolved in 0.1 N HCl, the cystine content of casein purified by heating with dil CH_3COOH was found to be 0.28% (uncorrected for moisture or ash). When the standard was cystine put through the same procedure of heat treatment with decolorizing agent, neutralizing, etc., that the casein went through, the cystine content of the casein was found to approximate 0.30% (uncorrected for moisture and ash). Crude casein and casein

different treatments were found to vary greatly. The loss of cystine in the ord-
 procedure of hydrolyzing decolorizing, etc., is mainly in the use of decolorizing

The formation of the dark colored humin, acid sol and acid insol, was checked
 uting the hydrolysis in a reducing atm such as furnished by SnCl_2 and TiCl_3
 ems more useful for the prevention of humin, since it can be pptd out by neu-
 on. The findings with Ti are slightly higher than without its use; but pend-
 study of the effects of high salt content, traces of Ti , etc., on the color develop-
 these findings are left open to future interpretation. There are 39 references
 literature. IX. Estimation of cysteine in the presence of glutathione. M.
 IVAN AND WALTER C. HESS. *U S Pub Health Repts* 46, 390-3(1931), cf
 5, 2470—Meldrum and Dixon (cf *C A* 24, 5315) found that the Sullivan
 for cysteine (cf *C A* 20, 2856) was markedly inhibited by the presence
 thione in the proportion of 9.0 mg of glutathione to 1.0 mg of cysteine. If
 ivan procedure detailed in a later paper (cf *C A* 23, 3941) had been followed
 author, they would have found that reduced glutathione in the proportion
 g to 1.0 mg of cysteine has no inhibiting effect on the estn of cysteine. To
 each soln and standard add 1 cc. of a freshly prepd aq soln of NaCN , shake,
 1.1 cc. of a freshly prepd 0.5% aq soln of sodium 1,2-naphthoquinone-4 sul-
 shake it (1 sec), add 5 cc. of 10-20% soln of anhyd Na_2SO_3 in 0.5 *N* NaOH ,
 1 and wait 1 min at about 20°. A reddish brown color appears. Then add
 a 2% soln of $\text{Na}_2\text{S}_2\text{O}_4$ in 0.5 *N* NaOH . The brown red color in the presence
 me is converted into a purer red. The reaction is given by no other compd
 not even glutathione or cysteine amine or even isocysteine. Even at the
 3 mg of glutathione to 1.0 mg of cysteine, the colorimetric reading of the 1.0
 the mixt was 8 when matched against 1.0 mg of cysteine similarly treated
 at 20. Higher glutathione content calls for more naphthoquinone. When to
 tathione-cysteine mixt 30 mg to 1 and to standard cysteine soln 1.0 mg in 5
 1 *N* HCl there were added 1 cc. of 1% aq NaCN and 1 cc. of a 1% soln. of the
 oquinone, followed by the regular Na_2SO_3 in the 0.5 *N* NaOH and then after
 color development by 1 cc. of the $\text{Na}_2\text{S}_2\text{O}_4$ in 0.5 *N* NaOH , no inhibition oc-
 For glutathione-cysteine mixt 100 to 1 the following procedure is recom-
 d. To 5 cc. of the mixt and standard add 1 cc. of 1% aq NaCN , 2 cc. of 1% aq
 a 1,2-naphthoquinone-4 sulfonate, followed by 5 cc. of 10% Na_2SO_3 in *N* NaOH
 after 30 min standing, by 1 cc. of a 2% soln of $\text{Na}_2\text{S}_2\text{O}_4$ in *N* NaOH . J. A. K.

semi-micro method for the rapid and complete analysis of human milk. MICHAEL
 ROBERTS AND A. LESRAGON. *Bull soc chim Biol* 12, 1195-1211(1930)—Detail-
 ren for the detn of total solids and Cl on 1 cc., casein and reducing value on 2
 ise of $\text{K}_2\text{Cr}_2\text{O}_7$ pptn, and fat and sugars on 1 cc. after saponification pptn and
 ith ether (Bang method). C. G. KING

determined carbon and glucidic carbon in normal urine. PAUL FLEURY AND
 AMBERT. *Bull soc chim. Biol* 12, 1275-68(1930).—See *C. A.* 25, 1862.

C. G. KING

ethods of determining bile salts in bile and duodenal juice. L. CUNY. *Bull*
m Biol 12, 1298-1318(1930), cf *C. A.* 25, 1862-3—A crit discussion is given of
 is in general use. C. G. KING

e quantitative limits of the Gerhardt and Legal reactions for the estimation of
 compounds. ADOLF FASCHERO. *Rev facultad cien quim (Univ La Plata)* 7,
 7-52(1930)—Legal's reaction as applied to urine is pos when acetone or diacetic
 exceeds 0.25 g per 1000 cc. The reaction of Gerhardt is pos. when diacetic acid
 0.2 g per 1000. B. S. LEVINE

ie preparation of a secretin concentrate. A. C. IVY, G. KLOSTER, G. E. DREW-
 o H. C. LUTHE. *Am J. Physiol* 95, 35-9(1930)—An improved method which
 ble, rapid and relatively inexpensive is described. The NaCl ppt. (*C. A.* 24,
 s extd with alc. at a concn. of 70-85%, which dissolves about 50 to 75% of the
 s. The exts. are evap to dryness over a water bath with addn. of water oc-
 ally to prevent inactivation of secretin. The residue is suspended in water and
 id for 10 to 12 hrs. Then NaOH is added to pH 5.3 to 5.7 when max floccula-
 cures. To the clear filtrate add enough 20% trichloroacetic acid to give 5%
 in the mix and let stand 4 to 12 hrs in a cold place, centrifuge and wash the
 ith anhyd, aldehyd-free acetone and ether. The av. effective dose of the ppt
 is 0.7 mg. It contains 1 dose of cholecystokinin for 7 to 10 doses of secretin
 t. From the isoelec. pptn contains secretin which can be extd by 0.4% HCl ,
 with $\text{CCl}_4\text{CO}_2\text{H}$ at 5% concn, sepd by the centrifuge and washed with acetone

and ether. This prepn is free from cholecystokinin. Its effective dose per dog is 0.61 mg. J. P. LYMAN

Cardiac output in normal men. I. STARR, JR., AND L. H. COLLINS, JR. *Am J Physiol* 96, 228-42(1931).—Expts on the fundamental assumptions of the improved Fick method for the detn of cardiac output were repeated with results that give greater confidence in the method. The cardiac output under various conditions was detd.

J. P. LYMAN

The measurement of the red-cell volume. I. PONDRA AND G. SASLOW. *J. Physiol* 70, 18-37(1930).—This is a colorimetric method requiring about 35 cc of blood and giving results agreeing closely with those calcd from cell measurements. Some values obtained by the method for man, rabbit and sheep, resp., are: 87, 58 and 30 μ^3 . The refractometric method of measuring red-cell vol led to erroneous results. J. P. L.

Measurement of red-cell volume. II. Alterations in cell volume in solutions of various tonicities. I. PONDRA AND G. SASLOW. *J. Physiol* 70, 169-81(1930).—The addn to rabbit blood and plasma of NaCl or glucose, of which Δ is approx the same as that of rabbit plasma (i. e., 1.2 g NaCl per 100 g water), does not alter the mean vol of the red cells. The cell vol increases in hypotonic solns of NaCl and plasma approx as would be expected on the assumptions that the red cells are simple osmometers contg 33% H₂O by vol. In hypotonic media, produced by the addition of KCl or glucose, the swelling is greater than can be accounted for in this way. The cell vol does not decrease in hypertonic media, produced by the addn of NaCl, KCl or glucose, but rather it steadily increases as time goes on, indicating that the added substances penetrate into the cell. Crenation is not necessarily associated with a decrease in red cell vol, but is regarded as a failure of the cell to maintain its shape. J. P. LYMAN

Recording of respiration of small animals. G. FINDERFELS. *J. Physiol* 70, 218-20(1930).—The app consists of a small covered box in which the animal is fastened, the animal breathes outside air through a breathing mask and an arrangement of valves. Changes in the vol. of the animal, due to breathing, are traced by the corresponding movements of a recording pen. J. P. LYMAN

The carbon dioxide of the mixed venous blood of men. W. P. HAMILTON, M. C. SPARBLIN AND H. G. SAAM, JR. *J. Physiol* 70, 244-52(1930).—There is no valid evidence in the exptl work reported in this paper that rebreathed air comes to equilibrium with mixed venous blood; hence there is doubt as to the justification for using any of the indirect applications of the principle of Fick in detg. the circulation rate in man.

J. P. LYMAN

A method of ultra-filtration in vivo. A. GRIGER. *J. Physiol* 71, 111-20(1931).—A specially constructed ultra filter is joined to an artery or vein; the blood circulates continuously through the app, and returns to the same blood vessel. The ultrafilter membranes are of collodion, and under negative pressure at their outer faces. The fluid in passing through the app retains its natural velocity, pressure and gas tension. A choice of membranes, filtrates contg. crystalloids only or a plasma which contains the whole or a part of its proteins may be obtained.

J. P. LYMAN

The measurement of the circulation rate in man by the acetylene method. A. GROLLMAN. *Proc. Physiol. Soc., J. Physiol* 70, xxxix(1930), cf. *C. A.* 24, 5050—C₂H₂ is sufficiently sol in blood and can be detd. very easily by alk. Hg(CN)₂, making it possible to complete a detn. of the cardiac output in less than 1 hr. The cardiac output of normal young individuals is a function of the surface area, hence it is predictable with a high degree of accuracy.

J. P. LYMAN

Development and present status of some problems and the goal of vital staining. JOSEF GICKLHORN. *Ergeb. Physiol* 31, 388-420(1931).—A review.

C. M. M.

A clinical method for the quantitative determination of pancreatic enzymes in duodenal contents. EDWARD HOLLANDER. *J. Lab. Clin. Med* 16, 400-5(1931).—The activity of the steapsin and trypsin present in the duodenal contents may be estd from a detn. of the length of time required for the neutralization of 1 cc of 0.1 N NaOH by the fatty or amino acids formed by the action of 1 cc of duodenal contents upon olive oil or gelatin, resp. The time required is normally 5 min. The amyllopsin activity may be estd from a detn. of the amt of substrate required to reduce 5 cc of Benedict's quant. sugar reagent after the action of 1 cc of duodenal contents upon sol starch. The reduction normally requires 19 cc or less. A decrease in the activity of the enzymes is usually assocd with disease of the gall bladder.

E. R. M.

Values in routine analysis. O. B. PRATT AND H. O. SWARTZOUT. *J. Lab. Clin. Med* 16, 471-5(1931).—Misleading values for urinary *pu* may be the result of dietary changes or medication before and during collection periods, or of detg. the *pu* of a single specimen rather than a collection over a 24 hr. period. For colorimetric

detns., deeply colored urines should be dild more than those that are slightly colored

E. R. MAIN

A simplification of the Osgood-Haskins hemoglobin method. EDWIN E OSGOOD AND HOWARD D HASKINS *J Lab Clin Med* 16, 482-6(1931); cf *C A* 17, 3517.—A soln. of acid hematin may be used as a standard in place of the permanent standards prep'd from morg salts

E. R. MAIN

A rapid method for determination of the sedimentation rate of the red cells with results in health and disease. HOWARD D HASKINS FRANK E TROTMAN, EDWIN E OSGOOD AND ALBERT MATHIEU *J Lab Clin Med* 16, 487-94(1931)—A method is described for use with oxalated blood in which the sedimentation rate may be expressed as the max settling in a 15 min period. The rate is normally 5 mm. but may become increased in pneumonia, acute inflammations and infections, acute rheumatic fever, infectious arthritis and malignancy

E. R. MAIN

The determination of blood proteins by a direct micro Kjeldahl method. ROGER S HUBBARD *J Lab Clin Med* 16, 500-3(1931)—A micro Kjeldahl method is described which is adapted to the detn of total N and the N of the protein fractions of the blood. The digested material is nesslerized directly. Pptn. of the Nessler's reagent is prevented by the use of Rochelle salt.

E. R. MAIN

A note on the Gunther-Greenberg method for determining inorganic phosphorus on the filtrate from calcium analysis. RUTH P BULTON *J Lab Clin Med* 16, 503-4 (1931), cf *C A* 23, 3942.—The morg P content of the filtrates from Ca analyses may be det'd by the Benedict-Thies method (*C A* 18, 3393), as well as by the Fiske-Subbarow method

E. R. MAIN

The use of the interferometer for serum protein and protein fraction determinations. WM J DRECKMANN *J Lab Clin Med* 16, 513-9(1931)—Methods are described for the detn of serum protein, fibrin and the albumin-globulin ratio by means of the interferometer.

E. R. MAIN

Determination of cholesterol in blood plasma and serum. J C FORBES *J Lab Clin Med* 16, 520-1(1931)—A modification of Bloor's method is described in which the cholesterol is ext'd. in the presence of Douci (a com prepn for softening water). No application of heat is required

E. R. MAIN

Phenol test for urinary albumin. WM R CLAPP AND BENJAMIN COHEN *New Engl. J. Med* 203, 1237-8(1930)—A method is described for the detection of albumin in the urine in which a ring test is obtained with a soln of CaH_2O_8 . The reagent is prep'd by treating a sat'd aq soln of CaH_2O_8 with sufficient glycerol to adjust the sp gr. to 1.045. Albumin present in the urine may be detected in a concn of 0.004%.

E. R. MAIN

Bilirubin liver-function test. I. Modification of the method. I. R. JANKELSON AND S. L. GARGILL. *New Engl. J. Med* 204, 547-9(1931)—A liver-function test may be carried out by detn of the concn of bilirubin in the blood at 5-min and 3-hr. intervals following intravenous injections of bilirubin. It is rapidly excreted through the liver in normal individuals but is definitely retained in diffuse disease of the liver.

E. R. MAIN

Separation of the male sex hormone from the female hormone, menformone. E. DINGEMANSE, J. FREUD, S. KBER, E. LAQUEUR, A. LUCHS AND A. W. P. MÜNCH. *Biochem. Z.* 231, 1-5(1931)—Menformone behaves somewhat like an acid and in a diphasic system, benzene-alk. 70% alc., it passes almost quantitatively into the latter phase. In extg the male hormone from male urine it is found to proceed more rapidly in a weakly alk. medium. The method of sepn. depends upon extg with benzene 2 portions of the urine, one made alk., the other acid

S. MORGULIS

The use of colloidal zirconic acid as a protein precipitant. JÓZSEF EPDŐS AND JÁNOS SÓRÓ *Biochem. Z.* 231, 6-12(1931)—Powd ZrCl_4 is rubbed up well with 4-5 times its wt. of anhyd AcOH and dried on a sand bath. The residue is digested for several hrs. with warm H_2O and the turbid soln is filtered. The filtrate is evap'd at 60° , the residue is moistened with H_2O and the evapn. repeated until no more AcOH fumes are given off. The dry residue is rubbed up with a little H_2O and is placed in a vacuum desiccator contg KOH sticks. The resulting $\text{Zr}(\text{OH})_4$ is practically free from Cl or AcOH and can be used in 5% soln. for deproteinizing serum plasma or protein solns.

S. MORGULIS

Methodical contributions. XIII. Determination of ferric ions, ferrous ions and of organically bound iron in biological material. L. PINCUSSEN AND W. ROMAN. *Biochem. Z.* 231, 54-8(1931)—A g. of fresh organ is quickly weighed and is finely rubbed up with sand in an ice-cold mortar. The material is washed twice with 5 cc. 20% H_2SO_4 into a centrifuge tube contg a little paraffin. In case of blood, 1 cc. of fresh blood

is measured under paraffin oil and treated with 5 cc 20% H_2SO_4 . The mixt is left for 3 hrs during which it is frequently stirred, and is sharply centrifuged. The supernatant fluid is quantitatively removed and is titrated with $TiCl_3$, KSCN being used. The ferrous ions are oxidized with HNO_3 and after driving off the HNO_3 excess the titration is repeated. The difference between the first and second titration measures the amt. of ferrous ions present. To det the organically bound Fe the residue is transferred from the centrifuge tube to a Kjeldahl flask by dissolving it carefully in HNO_3 . The material is ashed by boiling with $H_2O_2-H_2O$, the excess of reagents being driven off. The residue dissolved in dil H_2SO_4 is titrated with $TiCl_3$ giving the total ferric ion content and the organically bound Fe is calcd by difference. The $TiCl_3$ soln is prep'd by boiling 2 cc. of com $TiCl_3$ with 4 cc. conc'd HCl to drive off H_2S and dilg to 300 cc. The soln is kept in a container with a special automatic buret attachment (figured in the text) designed for titrating in a current of N_2 or CO_2 under complete exclusion of air. The $TiCl_3$ soln is standardized against a known Fe Cl_3 soln contg 10 mg Fe $^{+++}$ in 1 cc. To 0.1 cc. of this soln are added dil H_2SO_4 and a few drops of 15% KSCN and the mixt. is titrated until the color disappears. S. MOROGLIS.

Studies on cholesterol metabolism. I. A gravimetric method for the determination of free and combined cholesterol in small amounts of blood. RUDOLF MANCKE *Biochem Z* 231, 103-9 (1931).—The extn is carried out by pouring 2.5 cc. serum in a fine stream into a 50 cc. vol. flask contg 30 cc. of a mixt of 3 parts abs. alc. and 1 part abs. ether. Under const. stirring the flask is submerged in a boiling water bath and is removed as soon as the contents begin to boil. After cooling the vol is made up with the alc.-ether mixt. and the material is filtered. Twenty cc. of the filtrate (in 2 portions) is cautiously evap'd in a test tube, the temp. being gradually raised to 60° to boil off the alc. until there is only 2 cc. of soln. In case the material is turbid, it should be shaken with 10 cc. ether, filtered, washed and again evap'd to 2 cc. To the still hot soln. 1 cc. of hot 1% digitonin in 90% alc. is added and after 1 hr the ppt is collected in a filter tube (Jena filter G 154, No 1). The ppt is washed with ice-cooled acetone, and the filter is dried at 100° and weighed, 1/2 of the pptd comp'd being cholesterol. For the total cholesterol detn 10 cc. of the original filtrate in 2 portions is boiled down to 3 cc. in a 8-cc. round bottom flask provided with a reflux condenser. To this is added 3 drops 80% KOH, and the material is vapond 6 hrs. over a water bath. The alc. soap soln is neutralized with a few drops of alc. acid with gaseous HCl (phenolphthalein is used as indicator), mixed with 3 cc. abs. ether and the pptd KCl is filtered off. The ppt. is washed with abs. ether, the combined filtrates being condensed to 2 cc. The pptn. with digitonin and weighing are made as before. The difference between the 2 detns. gives the cholesterol esters. S. MOROGLIS.

Method of metabolism studies in tissue cultures. I. Respiratory measurements in tissue cultures. ROLF MEYER *Biochem Z* 231, 247-52 (1931).—The Carrel tissue culture flask is adapted for the detn of respiration by means of a specially constructed manometer inserted into the flask. II. Weight determination on single tissue cultures. Increase in weight and area. *Ibid* 233 9.—A method is discussed for measuring the wt. of a tissue culture after the removal of the plasma coagulum. These studies show that the increase in surface must be due partly to cellular migration. In rapidly growing fibroblast cultures the daily lactic acid production is 2 1/2 times the dry wt. and in slowly growing cultures 3 times the dry wt. S. MOROGLIS.

Determination of urobilin and urobilinogen with the Zeiss step photometer. L. HEILMEYER AND W. KREBS *Biochem Z* 231, 393-8 (1931).—The Zeiss step photometer has been employed to det urobilin and urobilinogen in urine and feces according to Terwen's procedure. This yielded greater precision, especially with small urobilin concns., and dispenses with the need for a comparison soln. which also saves much time in carrying out the detn. S. MOROGLIS.

An apparatus for the rapid, accurate analysis of the gases in a respiration chamber. THORNE M. CARPENTER, ROBERT C. LEE AND ANNA E. FINNERTY. *Miss Arch Landw. Abt. B. Tierernahr. Tierwcht* 4, 1-26 (1930).—The app. described by Carpenter, Fox and Sereque (*C. A.* 23, 4718) has been provided with a bellows to circulate the pyrogallie acid soln., thus reducing the time required for the analysis. The app. is described, and directions are given for its use. W. GORNOV ROSE.

Micro chloride determination in blood. J. A. FLEET KOK. *Arch. nederl. physiol.* 16, 132-5 (1931).—Dil the blood with an equal vol. of H_2O ; to 0.2 cc. add 1 cc. H_2O plus 0.4 cc. of 10% sulfosalicylic acid, shake thoroughly and centrifuge. To 0.8 cc. of filtrate add 1 drop of Na nitroprusside and then titrate with 1/10 N $Hg(NO_3)_2$. A correction is made according to the tables of Kolthoff and Bak for the quantity of $Hg(NO_3)_2$ in excess necessary for the formation of the turbidity. M. H. SOLTZ.

Determination of calcium in blood. CH O GUILLAUMIN *J pharm chim* [8], 13, 65-76 (1931) — Detn is made of total Ca in plasma or serum regardless of the mode of physico-chem combinations of Ca in the plasma. Precautions in taking samples are discussed, and the technic of detg Ca in plasma is given in detail. S. W.

Dichlorofluorescein as an adsorption indicator for the estimation of chlorides in the blood. A F OSTERBERG *Proc Staff Meetings Mayo Clinic* 5, 300 (1930) — Plasma or serum, 2 cc, is added, with shaking, to 7 cc acetone contained in a centrifuge tube graduated a 110 cc. The vol is made accurately to 10 cc, the tube is then stoppered, shaken well and the blood proteins are thrown down in the centrifuge. Five cc. of the acetone soln corresponding to 1 cc of plasma is pipetted into a small Erlenmeyer flask, and to this is added 0.25 cc of the indicator soln prep'd as described by Kolthoff, Lauer and Sunde. The soln is then titrated with N/3546 soln of AgNO₃ until the end point is reached. This manifests itself by the pptn of brick-red silver dichlorofluorescein on the silver halide ppt. One cc AgNO₃ soln is equiv to 1 mg Cl. This method checks closely with the Wilson and Ball procedure, the maximal difference being 1.6%.

R C WILLSON

Determination of traces of Hg (Stock, Lux) 7

Centrifugal blood-separating apparatus for serum plants. CECIL E MITCHELL
S 1,707,876, March 24. Structural features

C—BACTERIOLOGY

CHARLES B MORREY

Variability of the diphtheria bacillus. M A KUSHNARYEV *Z Immunitäts* 68, 210-7 (1930) — The addn of minute quantities of MnSO₄, ZnSO₄, LiCl and NaNH₂HPO₄ to the medium enhances the growth of diphtheria bacilli and causes a dissociation of the original strain into toxigenic and atoxigenic forms, as well as the production of coccus-like forms.

H EAGLE

Microflora of processed cheese. J CSISZÁR *Kísérlet Kozlemények* 33, 391-401 (1930) — Processed cheese contains chiefly lactic acid bacteria (streptococci and lactobacilli) and very often anaerobic butyric acid-producing organisms, accompanied by aerobic spore-forming bacilli (*B. mesentericus* and *subtilis*), indifferent, heat resistant cocci and bacilli, molds and yeasts and sometimes *Act. odorifera*. The variety of the flora and the no. of bacteria depend on the raw cheese, on the time and temp. of melting and finally on the age of the processed cheese. The germ content averaged 19,500,000; the highest no. was 310,400,000 per g of cheese. Two samples were found to be sterile. Heating to 65-80° for 5-25 min decreased the no. of germs by 99.9-100%; processed cheese is thus quasi-pasteurized. The germ content of ready-made cheeses increased from 5-27,000 to 1,800,000-258,900,000 organisms in 40 days.

S S DE FINÁLY

Bactericidal action of "Rohchloramin" and "Streuchloramin." N KERRLER *Médegyszerész Kutatások* 2, 319-24 (1929) — The action is about the same as that of HgCl₂. A 0.1% soln of "Rohchloramin" proved to be effective. "Streuchloramin" killed in less than 5 min: *B. abortus* inf. Bang, *B. gallinarum*, *B. coli*, *B. cholerae* gallinarum, *B. erysipelatis* suis and *B. suispestifer*; within 20 min. it killed *Staphylococcus pyog. aureus*. Spores of *B. anthracis* were killed in 60 min.

S. S. DE FINÁLY

Observations on the value of a copper sulfate tellurite medium for the isolation of diphtheria bacilli. V D ALLISON *Brit J. Exptl Path* 11, 244-8 (1930) — Tables are given to compare the results obtained in the isolation of *Corynebacterium diphtheriae* and diphtheroid bacilli from nasal, throat and ear swabs, with Douglas's medium and the copper sulfate tellurite serum agar medium of V. D Allison and Ayling (*C. A.* 23, 3949). The results show a well-marked difference in the percentage of organisms isolated on the 2 media, the differences in all cases being in favor of the CuSO₄ medium. This medium consists of 2% nutrient agar to which is added 10% of trypsinized horse serum, 0.02% of K tellurite and 0.05% of CuSO₄. The trypsinized serum enhances the growth of the diphtheria bacilli, while the K tellurite darkens the colonies and gives them a characteristic appearance. The growth of staphylococci, streptococci, pneumococci and *Micrococcus catarrhalis* is completely prevented by the presence in the medium of the small quantity of CuSO₄, which also inhibits completely the spread of *B. proteus*. The nutrient agar should not be more acid than pH 8.0 previous to the addn of trypsinized tellurite serum and CuSO₄, as it is impossible to measure the pH with indicators after the CuSO₄ has been added.

HARRIET F. HOLMES

Virulence of hemolytic streptococci. II. The influence of oxygen on the maintenance of virulence in broth cultures. E W TOWN *Brit J. Exptl Path* 11, 469-79

(1930), cf C A 25, 1273—Virulent hemolytic streptococci maintain their virulence after subcultivation in aerated broth and they change to the most attenuated form after subculture in anaerobic broth. The effect of aeration is increased if peroxide accumulation is prevented by catalase. Most attenuated cultures are not reverts to the virulent form by subcultivation in aerated catalase broth. III. The influence of oxygen on the restoration of virulence to most attenuated cultures. *Ibid* 487-8.—The increase of virulence which occurs when most attenuated cultures of hemolytic streptococci are subcultured in 50% normal horse serum is dependent on the O tension of the serum. The optimum condition for increasing virulence is the highest O tension which can be obtained without peroxide formation. The effect of normal serum on virulence appears to depend on its O-carrying capacity. HARRIET F. HOLMES

Investigations of the biologic, serologic and colloid-chemical behavior of the paratyphus bacillus under changed living conditions. W. D. SUTHER Arch Hyg 104, 239-54 (1930)—Continued cultivation of the paratyphus bacillus upon agar contg 0.15% phenol or 3-5% NaCl appears to produce changes in colony form and reductions in acid flocculability, cataphoresis, tendency to be pptd by $HgCl_2$ and capillarity. E. R. MAIN

Bacteriology of mucous and bitter tasting milk (CSISZ) 12. Syntheses of antiseptic derivatives of indan) 3 done WALKER *et al* 10. The correlation of the oxidation of certain phenols and of dimethyl p-phenoxydimethylamine by bacterial suspensions (HAPPOLD) 10.

D-BOTANY

THOMAS G. PHILLIPS

The pigment of *Vellula spirans* and *Fiona maritima*. BENJAMIN KROFF *Biol Bull* 60, 120-2 (1931)—The pigment consists, apparently, of a complex protein combination. There are as yet, no data as to its true structure. FREDERICK G. GERMUTH

Recent advances in science plant physiology. J. C. WALTER STILES *Science Progress* 25, 510-4 (1931). A review of recent work on photosynthesis. J. S. H.

The composition of some green forage and fertilizing plants. T. NASELHOFF, T. HAUN AND W. ELBERT *Landw Vers Sta* 110, 268-83 (1930)—Analyses of several varieties of clover show that the amt of N in the part of the plant above ground is much higher than that in the roots. JOHN R. HILL

The ether extract of white leaves of cabbage. III. The unsaponifiable matters. JUNICHI OKAMI *J Agr Chem Soc Japan* 6, 773-82 (1930), cf C A 25, 984—The oil of white leaves of cabbage was saponified and extracted with ether. The ether ext was washed with water. The waxy substance appeared between the ether and water layers. It was crystd from acetone, m. 81.5-2°, sol in $CHCl_3$. It shows no sterol reactions. Analysis of the cryst substance and of its oxide were made. It is presumed to be palmitone. From the mother liquid of palmitone an amorphous ppt was obtained. It m. 68° and is sol in alc. and acetone. It coincides with α -hentriacontane. O pre-(condensation) (reduction)

sumes the following changes. Palmitic acid $\xrightarrow{\text{condensation}}$ palmitone $\xrightarrow{\text{reduction}}$ hentriacontane. From the ether ext of the unsaponifiable matters a kind of phyto-sterol was isolated. Carotene and xanthophyll appear to be present. Y. KIHARA

The saponin of soy bean. II. YUSUKE SUMIKI *J Agr Chem Soc Japan* 6, 783-90, *Bull Agr Chem Soc Japan* 6, 49-51 (1930), cf C A 24, 3813—The Na salt of the saponin (decompt at 259-60°) was prepd as hexagonal plates by neutralization with NaOH from a 70% alc. soln of free saponin. The hemolytic power of soy-bean saponin was observed in a concn of 1/2000 and that of the Na salt in a concn of 1/50,000. Merck's saponin (from the root of *Saponaria*) shows hemolytic power in a concn of 1/50,000. The toxic power of soy bean saponin (free and the Na salt) is very weak for pigeons while Merck's saponin is very toxic. Y. KIHARA

The chlorophyll content of the leaves of rice plant. KENKICHI SATO *J Sci Agr Soc (Japan)* No 326, 24-36 (1931)—Willstätter's colorimetric method for the detn of chlorophyll was used. The chlorophyll content does not vary greatly with the species. It increased rapidly in the earlier stage of the growth, reached a max in July and then decreased gradually. The greater the growth rate of the plant the greater the pigment content. The amt of fertilizers seems to be approx proportional to the chlorophyll content. Deficiency of N significantly inhibits the formation of chlorophyll, while that of K or P inhibits it less or negligibly. Y. KIHARA

The relation of the various physiological changes of plants and their vegetation periods to pigment formation. II. Assimilation rate and anthocyanin synthesis in

Abutilon avicennae. HIROSI KOSAKA. *J. Dept. Agr. Kyushu Imp. Univ.* 3, No. 2, 29-45 (1931), cf. *C. A.* 24, 2159-60.—Anthocyanin synthesis in stems and leaves of *Abutilon avicennae* is parallel to the content of assimilation products and in inverse relation to the rate of growth. H. R. KRAYBILL.

Culture and nutritional physiology of the genus *Pilobolus*. EGON BERKA. *Sitzb. Akad. Wiss. Wien, Abt. I*, 139, 355-71 (1930).—*Pilobolus klivi* and *Sphaerosporus* can be cultivated easily upon horse-manure-decoction agar. Despite favorable culture conditions and frequent transplantations some of the mold cultures die, the reasons for this are not known. Content of the horse-manure-decoction-agar favors the mold growth. Metabolic expts. upon pure cultures proved that the materials used as sources for N are divisible into (a) favorable: peptone, albumin, leucine, asparagine, (b) barely sufficient: casein, glycocoll, alanine, glutamine, (c) insufficient: urea, glucosamine, NH₄ salts. Materials used as sources of C are divisible into (a) those favoring development of growth to the point of sporangia formation: xylan, gum arabic, arabinose, galactose, (b) those impeding or arresting the mycelial development: starches, inulin, fructose, glucose, saccharose, lactose, mannitol. Boiled wheat straw, with or without peptone, or its exts. forms a favorable medium, pectins, from the cortex of apples, are not favorable. Nuclein and Liebig's beef ext. favor the mold development. This would indicate that nucleic acid contains pentosan and that split products of nuclein occur in the beef ext. The pos. phloroglucinol-HCl reaction appears to substantiate such an assumption. *P. sphaerosporus* differs from *P. klivi* Tiegh only in its manner of germination, but not in its process of metabolism. It is possible that *P. heterosporus* Palla, found only on cow manure, behaves differently as regards its nutritional physiology. B. S. LEVINE.

A preliminary report upon the results of field experiments with substituted products for sugar beets. J. PÄTZER. *Listy Cukrovár.* 49, 366 (1931).—A criticism is given of the study in which 40 expts. with sugar beets are compared with 30 expts. with edible carrots. P. considers the statistical analysis used as inadequate. FRANK MARESH.

The evaluation of the roots of fodder beets for eugenic purposes. K. KOČNAR, J. HAMPL AND VL. ŠMERDA. *Sborník Čsl. Akad. Zemedělstv.* 6, 95 (1930), *Listy Cukrovár.* 49, No. 26, Rozhledy 22.—Chem. analysis was found of little value in studying eugenic expts. with fodder beets, polarization did not give the true sucrose content, for inversion occurs readily, and the content of invertase is increased, the refractometric detn. can be used only for roots of the same dimensions and variety in order to have comparable results. The beets are mashed, weighed and digested with hot water; the w is then read and the sucrose content computed. The margin of error is $\pm 1\%$, but the results are more accurate than with a polarimeter. The distribution of sol. nutritional elements varies in the same species and is altered by the shape of the root, removal of greens, depth of root in the ground and accessory roots. The instructions for cutting cylinders for sugar detns. from sugar beets do not hold for fodder beets, and a trial detn. of the distribution of sugar must be made on each variety before making dependable analyses for the whole plot. The superficial strata were richer in sugar than the inner ones. FRANK MARESH.

Formation of the anthocyanic pigment in the etiolated plants of buckwheat and wheat. SR JONESCO. *Compt. rend.* 192, 438-43 (1931).—Using the same method of extn. of pigment (*C. A.* 17, 792, 22, 102) J. studied exts. of buckwheat and wheat to det. whether they contain a substance, similar to that obtained from *Ampelopsis heterophylla*, which can be transformed into an anthocyanin. From buckwheat, a chromogen is obtained as a yellow powder which, when hydrolyzed, forms galactose, when the chromogen is oxidized or dissolved in alc. and heated with HCl, a dense-red cryst. pigment is obtained which shows the reactions characteristic of a natural anthocyanidine. From wheat, the chromogen is a yellow powder, sol. in water, not sol. in alc., has a sweet odor, reduces Fehling's soln. and hydrolyzes to form arabinose. Oxidation of the chromogen with BaO₂ or MnO₂ in H₂SO₄ soln. forms a violet red pigment characteristic of the natural pigment of wheat plants. N. M. NAYLOR.

The presence of allanto-nase in a number of fungi. A. BRUNEL. *Compt. rend.* 192, 442-4 (1931).—In 67 different fungi, a study was made of the enzyme which changes allantoin to allantoic acid. The allantoic acid is recognized by its products of hydrolysis, glyoxylic acid and uric acid; or by the formation of the Ag allantoate. The enzyme action is tested as follows. The macerated vegetable tissue is extd. with glycerol and water. The ext. is added to a medium contg. allantoin, basic NH₄ carbonate and a few drops of chloroform. After 12 hrs. of digestion, the mixt. is made acid and heated to 100°. The product is tested for glyoxylic acid by noting a red coloration in the

presence of phenylhydrazine and K_2FeO_4 , it is tested for uric by the addition of methyleanthranil which forms a ppt. of xanthine.

The stimulation of electrolysis. II. Suggestions as to the nature of accumulation in Valonia. W. J. V. OOSTERHOUT. *J. Gen. Physiol.* 14, 285-331 (1931).—In a living cell like that of the marine *V. natans*, K^+ probably enters chiefly as KOH , the thermodynamic potential of which is greater outside than within. The energy necessary to produce the accumulation of K^+ inside the cell is furnished by metabolism, whereby a weak org. acid, H_2A , is produced in the cell. H_2A is exchanged for HCl of the external soln. (or A^- for Cl^-) and this results in a continual maintenance of HCl at p_H 5.5, whereas the p_H of the sea water is about 8. KOH dissolves in to neutralize the acid within the cell, and hence KCl accumulates in the cell sap. Na^+ enters more slowly and does not reach an internal concn. as high as that of K^+ . Cell penetration is chiefly in mol. form. The system is not in equil. but depends upon the continual expenditure of energy to produce acid within the cell. III. Behavior of sodium, potassium and ammonium in Valonia. A. G. JACQUES AND W. J. V. OOSTERHOUT. *Jol* 301-14.—When 0.001 M NH_4Cl is added to sea water containing cells of *V. natans*, undissolved NH_3 or NH_4OH penetrates rapidly, increasing the internal p_H value so that the thermodynamic potential of KOH becomes greater inside than outside the cell. K^+ therefore leaves the cell, but $NaOH$ continues to pass in because its outside thermodynamic potential is greater than the inside potential. NH_4Cl accumulates in the cell, reaching a much higher concn. inside than in the outside soln. After entering, NH_3 probably combines with a weak org. acid produced in the cell. The action of this acid is exchanged for Cl^- of the sea water (or the org. acid is exchanged for HCl). Methods are described.

Aluminum content of plants, chiefly food plants. GABRIEL BERGMAN AND GEORGE L. LEE. *Genes* and 192, 523-9 (1931).—The plant is dried and Al detd. gravimetrically as Al phosphate, the details of analysis not being given. The Al content passes through a max. during the first period of the growth of the plants. Of the values given, expressed as mg. Al per kg. dry weight, are young carrot 30, old carrot 22, potatoes 78, onion 95, apple 13, tomato 25, apricot 16, peach 14, maize 15, white rice 14, coffee 49, beans 74, tea 55, Ceylon tea 455, rhubarb 165. Green leaves are rich in Al. Al may play some part in plant synthesis.

A study of twenty-three acyl alcohols on growth of *Lupinus* seedlings. DAVID I. MACINTYRE. *Am. J. Botany* 17, 572-8 (1930), cf. *C. A.* 24, 5884.—Twenty-three monomeric acyl alcohols were found to have different degrees of toxicity for *Lupinus albus* seedlings. A definite difference in toxicity was found between the primary, secondary and tertiary acyl alcohols, the primary being the most toxic and the tertiary being the least toxic. Differences in toxicity between the individual members of each group were also noted. Contributions of the various acyl alcohols exhibited synergistic effects. The results of the investigation emphasize the importance of employing pure chemicals in pharmacological research.

The use of carbon dioxide for prolonging the life of cut flowers, with special reference to roses. KENNEDY C. THORNTON. *Am. J. Botany* 17, 614-2 (1930), cf. *C. A.* 25, 5041.— CO_2 treatment was effective in prolonging the life of cut roses when they were stored for 3 to 7 days at 33° and at 50° in a concn. of CO_2 between 5 and 20%, when they were subjected to an alternated CO_2 -air treatment for 3, 8 or 14 days CO_2 retarded the opening of rose buds. This effect was more pronounced as the concn. of CO_2 was increased. CO_2 treatment appeared to be more effective for flowers in the bud stage than for those which were already open.

The hormone and climatic distribution of oil and starch in seeds in relation to the physical and chemical properties of both substances. JAMES B. MCNAUL. *Am. J. Botany* 17, 622 (1930), cf. *C. A.* 24, 4967.—Oil is more abundant than starch in seeds and in the seed embryos and albumens. In the seeds of temperate plant families starch is present less frequently than in the seeds of tropical plant families. The large proportion of oil in the seeds of tropical and temperate plants may indicate its function as a protection against rapid temperature changes in accordance with its heat conduct. which is lower than that of starch. The larger proportion of oil to starch in the seeds of temperate plants is in agreement with the greater fuel value of oil.

Sucrose and starch in potatoes treated with chemicals that break the rest period. F. E. DENNET. *Am. J. Botany* 17, 806-37 (1930), cf. *C. A.* 24, 399.—Freshly harvested potatoes were treated with CH_3Cl , CH_3OH , $NaSCN$ and thiourea, the contents of the chemicals being decreased from the epidermis by steps to form a graded series. Presences from the treated potatoes obtained from the various lots at a subsequent interval, usually 4-7 days before sprouting became visible, were compared with pieces from the

checks with reference to sugar content. Sucrose was found to be higher in the treated than in the check lots and to give a graded series of values corresponding to the series of contents of chemicals used in treating the potatoes. The reducing-sugar values did not form such a series, and no consistent effect of the treatments in either increasing or decreasing the reducing-sugar content was found. When samples were taken at intervals of 21, 48 and 72 hrs. after treatment it was found that the time after treatment at which the sucrose content of the treated lots became higher than that of the checks differed in different expts. from 21 to 72 hrs. Samples of entire tissue which had been dried, powdered and analyzed for starch showed that contents of chemicals favorable for breaking the dormancy of the sprouts caused decreases in the starch.

J. J. SKINNER

Light and permeability of protoplasm. V. V. LEFFSHIN. *Am. J. Botany* 17, 953-70 (1930).—The influence of light upon the permeability of protoplasm for aniline dyes was investigated. The permeability was measured according to the amt. of dye accumulated in the cell sap of *Elaeagnus* leaves, det'd by a colorimetric method especially worked out for *Elaeagnus*. The difference in absorption of dyes by plant cells in light and in the dark is caused by the change of permeability of protoplasm. The absorptive power of the cell sap and the adsorption of dye on the cell walls or in dead protoplasm are not changed by light. dead cells absorb dyes with the same speed in light as in the dark. The max. amt. of dye absorbed by the cell does not depend upon the illumination. The solns. of dyes fade in light, and must be changed during the expt. or flow through the vessels contg. the object. If this condition is fulfilled, the permeability of protoplasm is found to be greater in light to all dyes which penetrate protoplasm and do not absorb violet rays. The increase of permeability is observed only in the leaves or their parts which are illuminated. The increase of permeability does not spread from the cells affected by light even to neighboring cells.

J. J. SKINNER

The swelling of citrus fruits. FLOWARD S. REED. *Am. J. Botany* 17, 971-82 (1930).—The structure of citrus fruits presents certain unique problems pertaining to water absorption and conductance since the fibrovascular system is mainly restricted to the mesocarp. Orange and lemon fruits absorbed water and various solns. with resulting increase in vol. Both acid and base were absorbed, but with lemons the latter produced somewhat greater swelling than the acid. Oranges in CaSO_4 solns. swelled promptly and showed no reversibility. Solns. of compds. which coagulate pectins caused incipient swelling, but it was followed by shrinkage. The translocation of liquids appears to depend upon their passage through the layers of hydrophilic colloids on the walls of the mesocarp cells and only to a small degree upon the participation of living cells.

J. J. SKINNER

A microchemical study of soy beans during germination. FLOYD W. VON OILEY. *Am. J. Botany* 18, 30-49 (1931).—The cotyledons of mature Manchou soy-bean seeds contain a large amt. of protein oil, some nonreducing sugar and a small quantity of starch. Transformation and translocation of these reserve foods were followed microchemically during the germination and growth of soy-bean seedlings in darkness. During germination the first changes found were the appearance of reducing sugar, an increase of starch in the hypocotyl, appearance of starch in the root cap and an increase of starch in the cotyledons. During the first 3 days of germination there was a large accumulation of starch in the apex of the hypocotyl, and of reducing sugar in the base of the hypocotyl and root. The amt. of starch in the cotyledons increased until the 5th day and then remained nearly const. until the 9th day, after which there was a rapid decrease. It disappeared from the palisade last. The starch disappeared from the different parts of the seedling in the following order: root, hypocotyl, epicotyl and cotyledons. Reducing sugar disappeared from the seedling in the same order, but 2 days later, except in the cotyledons where only a slight test for it was obtained. Non-reducing sugar was not detected in the hypocotyl, plumule and cotyledons after the 3rd, 4th and 7th days, respectively. Asparagine appeared in the hypocotyl on the 3rd day. It gradually increased during the development of the seedling, until on the 29th day it was abundant in the hypocotyl, and fair amts. occurred in the base of the root and epicotyl. Asparagine was not detected in the plumule or cotyledons or at the root tip. The depletion of the oil of the cotyledons began at the base and progressed toward the opposite end. There was a gradual change of the organically bound P and Mg of the cotyledons to the inorganically bound. Although both inorganically bound P and Mg moved quite rapidly toward the growing points, they gradually increased in amt. in the cotyledons. K moved out of the cotyledons more rapidly than did P and Mg.

J. J. SKINNER

Studies on the sensitivity of *Mimosa pudica*. II. The effect of animal anesthetics

and certain other compounds upon sea-anemone sensitivity. RAYMOND H. WALLACE. *Am. J. Botany* 18, 215-35 (1931).—In expts. with *Miura's padina* cones, of ether vapor from 15 to 25% prevented the movement of leaflets and petioles in 10-45 min. after exposure. The leaflets regained their sensitivity within a few min. after removal from the ether vapor, but a period of 2 or more hrs. was required for the primary petioles to become normal. Chloroform in concns. of 2.5% or greater was lethal and lower concns. caused injury after long exposures. A concn. of 3% ethyl alcohol reduced the angle of movement of the primary petiole less than 4%. Chloroform vapor of a wide range of concns. induced a strong chemotactic response in the leaflets and petioles of *Miura*. CCl_4 induced a reaction very similar to that of chloroform. The physiologically significant range of concns. was similar to that of chloroform. *Miura's* plants retained their sensitivity 1-4 hrs. in very high concns. of N_2O and ethylene. Concns. of approx. 100% acetone made the plants insensitive, while 95% made them only partially insensitive, and 85% seemed to have little or no effect. Repeated asphyxiation with ether, or repeated exposures to N_2O , ethylene and acetone had no harmful effects on the plants. A 10% concn. of CO_2 made the plants completely insensitive in 10 min. Longer exposures caused serious injury or death. MeOH and EtOH seem to have only a slight effect upon sensitivity; the former decreases it slightly and the latter increases it slightly. The different compounds tested affected the position of the petiole in various ways; some caused an elevation, some a depression and some seemed to be without effect.

J. J. SKINNER

The natural coloring matter of raw silk fiber of the domestic cocoon. II. Xanthophyll of mulberry leaves as a source of yellow cocoon xanthophyll. MASAKI OKU. *J. Agr. Chem. Soc. Japan* 6, 1103-8 (1932), *Bull. Agr. Chem. Soc. Japan* 6, 104-6 (1932), & *C. A.* 24, 3884.—Fresh mulberry leaves were finely chopped and soaked in 5 l. of 90% EtOH and digested with 60% EtOH under 15 lb. pressure for $\frac{1}{2}$ hr. The treatment was repeated. All the alc. ext. of the chloroplast pigments was cooled, and ether was added. After evap. and apcn. with KOH methanol soln. the ether soln. was washed with water to remove K chlorophyllates. The ether soln. was dehydrated by anhyd. Na_2SO_4 and evapd. to dryness. 60 cc. of petroleum ether (b. 40-60°) was added. Xanthophyll was pptd. It was recrystd. from boiling MeOH . The yield was 0.165 g. from 1 kg. of leaves. It was identified with xanthophyll (m. 175°) from the domestic yellow cocoon. The origin of xanthophyll of the domestic yellow cocoon may be that of the mulberry leaves.

Y. KIBARA

Botany, chemistry and pharmacodynamics of *Toddia aculeata* (LOBSTERY, HESSE) 17. Plant pigments (KARREN, # 6) 12.

OSLOW, MORRIS W. The Principles of Plant Biochemistry. Part I. New York: The Macmillan Co. 330 pp. About \$4.50.

E-NUTRITION

PHILIP B. RAYE

Carotene from lettuce and its relation to vitamin A. H. S. OLCOVICH AND H. A. MATHILL. *Proc. Soc. Exptl. Biol. Med.* 28, 243-1 (1933).—Carotene was obtained from a MeOH ext. of the unsaponifiable fraction of the lipids of lettuce. After recrystn. from petroleum ether the m. p. was 179.5-180°. The yield from approx. 150 kg. of lettuce was 20 mg. On long standing at room temp. or on heating for 24 hrs. at 105° the hexagonal crystals assumed an isometric form, they bleached without change in wt. and lost their physikal. activity. Carotene is best preserved in Et laurate concn. 0.1% hydrogenase. Vitamin-A-deficient rats were cured of xerophthalmia and their growth increment was restored by daily doses of 0.005 mg. of carotene. C. V. B.

Hematopoietic function in avitaminosis. IV. Further studies of vitamin A deficiency. BARNETT SURE, M. C. KIRK AND DOROTHY J. WALKER. *Proc. Soc. Exptl. Biol. Med.* 28, 493-6 (1933), & *C. A.* 23, 4724.—The red-cell count, the hemoglobin concn. and the differential leucocyte count of the rat remained unchanged during the development of, and the recovery from, the severe ophthalmia and loss of body wt. associated with vitamin A deficiency. V. Vitamin D deficiency. BARNETT SURE AND M. C. KIRK. *Ibid.* 49-6.—The hematopoietic function was unaffected by vitamin D deficiency induced in the rat by feeding the Steenbock Black ration No. 2945.

C. V. BAILEY

Mineral metabolism in infancy. III. The substitution of cow milk for human milk in infant feeding. S. V. TYLOR. *Lancet Med. J.* 31, 273-84 (1933), & *C. A.* 24, 4829.—The changes in retention value and mode of excretion of the mineral elements

when whole cow milk is substituted for human milk have been studied by metabolic expts. The proportion of the total mineral intake retained is reduced although the actual quantity of bone-forming elements is substantially increased while that of Fe is diminished. On cow milk the intake of Ca, Mg and P is 4 times greater than on human and the increased retention is due to the larger quantities available for adsorption and fixation by the growing tissues. The fat intake is less effectively utilized than in natural feeding because of increased soap formation in the intestines. The concn of fat and mineral constituents is increased, the rise in the percentage of fat being due to combined fatty acids. The fecal wt and the output of phosphoric acid by the urine are greater than on human milk. These changes depend mainly on the differences in mineral compn of the 2 milks. RACHEL BROWN

The local calcification of tissue after subcutaneous administration of irradiated ergosterol. TH. V. BRAND AND F. HOLTZ. *Z. physiol. Chem.* 195, 241-7 (1931).—Daily subcutaneous injections of soy-bean oil and of non-irradiated ergosterol over a period of several weeks caused abscesses at the site of injection but no local calcification in the rats. Irradiated ergosterol, on the other hand, produced local calcification in the subcutaneous tissue. Calcification in the kidneys was much less with subcutaneous than with oral administration of the same dosage. The same effect was obtained with preps in which vitamin D had been destroyed by heating. It is due to the specific action of a calcemosis factor previously discussed (cf. Holtz and Schreiber, *C. A.* 24, 5799).

Experimental production of xerophthalmia and keratomalacia by the feeding of bread. I. ABELIN. *Arch. exp. Path. Pharmacol.* 155, 46-50 (1930). H. EAGLE

The isolation and the chemical and physical nature of vitamin A. HANS SEEL AND F. DANNMEYER. *Strahlentherapie* 39, 449-51 (1931).—The authors isolated vitamin A from the whale-liver oil, which has 10 times more vitamin A than cod liver oil and compared its physical and chemical properties with those of a substance obtained by careful oxidation of cholesterol. On careful oxidation of cholesterol, a different product is obtained from the well-known crystallized oxidation products of cholesterol, which showed a distinct vitamin A action. The identity of the natural and the synthetic vitamin A was demonstrated by the authors in the following way: (1) Both are amorphous masses of honey-like consistency. (2) Both will give the usual cholesterol reaction. (3) Both will give the Rosenheim and Drummond reaction of vitamin A with SbCl_3 in a very intense degree. (4) They show almost identical absorption of ultraviolet light. (5) The natural vitamin A shows an antixerophthalmic and a growth-promoting effect in a dose of 0.0001 mg. The synthetic product shows similar effect, but only up to a dose of 0.1 mg. This would indicate that on oxidation of cholesterol only a part of the cholesterol has been transformed into vitamin A. There is no doubt that vitamin A also possesses a sterol character. F. R. GREENDAUM

Alcohol solubility of the antidermatitis, more heat-stable vitamin B_2 constituent of the vitamin B complex. HARRIETTE CHICK AND ALICE MARY COPPING. *Biochem. J.* 24, 1744-7 (1930).—The addn of 56% alc. to a soln of vitamin B_2 at pH 1.5 (prepd from brewer's yeast) gave an inactive ppt, and the filtrate was $\frac{1}{2}$ as active as the original material. With 70% alc., a filtrate was obtained that was completely inactive. At pH 3.2 the filtrate was inactive when 56% alc. was used. BENJAMIN HARROW

Egg white as a source of the antidermatitis vitamin B_2 . HARRIETTE CHICK, ALICE MARY COPPING AND MARGARET HONORA ROSCOE. *Biochem. J.* 24, 1748-53 (1930).—Egg white (from hen eggs) contains no vitamin B_1 but is rich in vitamin B_2 . By removing the coagulable proteins, a rich ext. of vitamin B_2 may be obtained. BENJAMIN HARROW

Distribution of the vitamin B complex. I. Leafy vegetables. MARGARET HONORA ROSCOE. *Biochem. J.* 24, 1754-63 (1930).—Watercress, lettuce, spinach and cabbage (dry weights) have a content of B_1 and B_2 about $\frac{1}{4}$ that of dried brewer's yeast. Their content of B_1 is lower than that of wheat germ or ox-liver and higher than that of egg yolk or ox muscle. Their B_2 content is lower than that of ox liver, equal to that of milk, ox muscle or egg yolk and higher than that of the pulses or cereals. B_1 is more concd in dark-green leaves than in the paler ones. On the whole, the vegetables tested were richer in B_1 than in B_2 . BENJAMIN HARROW

Composite nature of the water-soluble vitamin B. III. Dietary factors in addition to the antineuritic vitamin B_1 and the antidermatitis vitamin B_2 . HARRIETTE CHICK AND ALICE MARY COPPING. *Biochem. J.* 24, 1764-79 (1930), cf. *C. A.* 24, 5801.—Yeast and watery yeast extracts contain, in addition to vitamins B_1 and B_2 , another dietary factor ("factor Y") which withstands prolonged heating in alkaline solution. B. II.

Biological values of proteins. I. A method for measuring the nitrogenous ex-

change of rats for the purpose of determining the biological value of proteins. HARRIETTE CHICE and MARGARET HONORA ROSCOE *Biochem J* 24, 1769-2(1930)—The av. daily nitrogenous balance sheet is held on a diet containing a definite amt. of the protein to be examined. Two adult male rats (about 250 g.) were placed in each cage, and detrits made of the total N ingested and of that excreted in the urine and feces, during a period of 4-5 days. To det. the bio. value of the protein, one det. the amt. quantity of nitrogen (N) which must be absorbed to compensate the daily nitrogenous expenditure (E) on a % free diet. The bio. value is $100 \times E/N$. II. Biological value of purified caseinogen and the influence of vitamin B₂ upon biological values, determined by the balance sheet method. MARGARET AVERIL ROAS FIKSEN *Ibid* 1794-1804—If purified casein was used and the B vitamins were added to the diet in the form of concentrates, the bio. value of casein was found to be 45—much lower than that found by previous workers. BENJAMIN HARKOW

Heart block in pigeons—curative factor. CYRIL WM. CARTER. *Biochem J* 24, 1811-9(1930)—Whole wheat and yeast contain a factor essential for the nutrition of the pigeon; the absence of which from polished rice leads to heart block. B. H.

A quantitative comparison of the curative activity of torulin (vitamin B₂) upon the adult pigeon and the adult white rat. HENRY W. KNEVERLEY, RUDOLF A. PETERS and VERA READERS. *Biochem J* 24, 1829-3(1930)—The B₂ factor for pigeons and rats is identical. The dose of B₂ required by each animal is about the same. B. H.

Curative activity of the antineuritic vitamin of rice. BARD C. P. JENSEN, HENRY W. KNEVERLEY, RUDOLF A. PETERS and VERA READERS. *Biochem J* 24, 1834-5(1930)—Using the Jensen and Donald rice vitamin crystals (*Proc. Acad. Wetensch. Amsterdam* 29, 1370(1929), cf. *C. A.* 21, 2159), the authors find the following activity (curative tests) per day dose of vitamin: pigeon 0.007 mg. by injection, 0.009 mg. by mouth, rat 0.005 mg. This is a higher activity than that obtained for torulin concentrates. BENJAMIN HARKOW

Assay of vitamin B₂. VERA READERS. *Biochem J* 24, 1827-5(1930), cf. *C. A.* 24, 402—Rats are deprived of the alkali labile vitamins B₁ and B₂ until polyneuritis occurs (3-4 weeks). The addition of B₂ cures polyneuritis, but does not restore weight. The animal remains in a weak condition with swollen red paws, sparse hair, loss of coordination, etc.—apparently because of lack of B₁. BENJAMIN HARKOW

Maintenance nutrition in the adult pigeon and its relation to torulin (vitamin B₂). I. CYRIL W. CARTER, HENRY W. KNEVERLEY and RUDOLF A. PETERS. *Biochem J* 24, 1832-43(1930)—B₂ from yeast will not give maintenance nutrition in pigeons unless supplemented by some undefined factor. II. *Ibid* 1844-51—Three B factors are necessary for nutrition in the pigeon. The factor which supplements B₂ is present in 20% alc. ext. of the charcoal process for coxage B₂ (*C. A.* 24, 2771) and also in alkali treated malt. It is not identical with B₁, B₂ or B₆, and can be termed B₃. B. H.

Relation of hydrogen-ion concentration to the precipitation of purified torulin yeast vitamin B₂ by phosphotungstic acid. HENRY W. KNEVERLEY and RUDOLF A. PETERS. *Biochem J* 24, 1546-61(1930)—By the use of fractional pptn. by phosphotungstic acid at varying pH values, B₂ concentrate with an effective action at 0.012 mg. is obtained. BENJAMIN HARKOW

Reaction of antimony trichloride with cod-liver oil and its unsaponifiable fraction. ERNEST L. SMITH and ROBERT HALEY. *Biochem J* 24, 1942-51(1930)—The unsaponifiable fraction gives with SbCl₃ a blue color proportional to its content. This fraction can be completely oxid. with H₂O, AcOEt, CHCl₃ or light petroleum. BENJAMIN HARKOW

Some contributing factors to the degenerative diseases, with special consideration of the role of dental focal infections and seasonal toxins in defensive vitamins. WESTON A. PRICE. *Dental Cosmos* 72, 1049-60, 1119-31(1930)—A study of the relationship between focal infections, dental caries and the vitamin content of the diet. JOSEPH S. HEPBURN

Metabolism of carbohydrates. FERNAND H. BADANEL. *Dental Cosmos* 72, 49-56(1931)—A review with bibliography. JOSEPH S. HEPBURN

Diet and dentition. WALTER H. FORD. *Dental Cosmos* 72, 546-52(1931)—A review devoted chiefly to the influence of the vitamins on the teeth. JOSEPH S. HEPBURN

Studies in the rat of susceptibility to dental caries. I. Bacteriological and nutritional factors. THEOPHIL ROSECRANCE and MAXWELL KARNESMAN. *J. Dental Research* 11, 123-35(1931)—Three synthetic ratons, (a) low in Ca, (b) vitamin D-free, (c) vitamin C-free, were fed to young rats. Each raton was fed both with and without the addition of human oral lactobacilli. Dental caries did not develop in any of the expd.

animals, even though organisms similar to *Lactobacillus acidophilus* occur normally in the rat mouth. A bibliography of 26 references is appended. J. S. H.

Pathological changes in the teeth of rats produced by synthetic diets. THEODOR ROSEBURY AND MAXWELL KARSHAN. *J. Dental Research* 11, 137-48(1931).—A synthetic ration low in Ca and free from vitamin D, produced in rat teeth histological changes ascribable to deficient calcification. The odontogenic zone was unusually broad in the dentine of both molars and incisors, large calcified globules appeared in the body of the dentine. The odontoblasts, the pulps, the enamel of the molars and the fully formed enamel of the incisors appeared normal, but the youngest enamel of the incisors and the ameloblasts appeared deformed. The staining quality of the alveolar bone was affected. Changes, similar in certain respects to those of latent scurvy in guinea pig teeth, were produced in the pulps and dentine of rat molars by a synthetic vitamin C free ration, and by other synthetic diets in which a supply of vitamin C was given daily as orange juice, the sp. cause of these changes was not ascertained, but apparently was not a deficiency of vitamin C. JOSEPH S. HEPBURN.

Relation between vitamin C and some phases of reproduction in the guinea pig. MARIANNE GOETTSCH. *Am. J. Physiol.* 95, 64-70(1930).—Female guinea pigs, gradually deprived of vitamin C, continued the estrous rhythm until they began to lose wt. Male guinea pigs, deprived of vitamin C, probably as a result of general vigor, lose the ability to sire litters, although motile sperms are present. J. F. LYMAN.

Physiology of exercise. IV. Exercise and basal metabolism in dogs. A. H. STENHAUS AND T. A. JENKINS. *Am. J. Physiol.* 95, 202-9(1930).—Conclusion. Exercise is without effect on basal metabolism. The high degree of relaxation which is acquired by the trained animal may explain why it is possible to have muscle hypertrophy without an increased metabolic rate. J. F. LYMAN.

The physiology of vitamins. XIV. The effect of administration of large amounts of water on the time required for development of the anorexia characteristic of a deficiency of the vitamin-B complex. G. R. CORGILL, H. A. ROSENBERG AND J. ROGOFF. *Am. J. Physiol.* 95, 537-44(1930), cf. *C. A.* 24, 4537.—Large vols. of fluid given by mouth to dogs subsisting on vitamin B-free diets shortened by about one half the time of appearance of the anorexia characteristic of starvation for the vitamin B complex. The hypothesis that the symptoms of lack of vitamin B are due to an accumulation of abnormal intermediary metabolites is unsupported by these expts. J. F. LYMAN.

Metabolism in pregnancy. IV. The nitrogen metabolism. A. W. ROWE, DOROTHY GALLIVAN AND HELEN MATTHEWS. *Am. J. Physiol.* 95, 592-600(1930), cf. *C. A.* 18, 3644.—Urinary N during pregnancy is lowered, urea N diminishes, NH₄, uric acid and undetermined N compds. of the urine increase. In the first few months after delivery the same general picture persists, except that uric acid elimination returns to normal. Non protein N, urea and residual N of the blood are all lower during pregnancy than at an equal non pregnant nutritional level. J. F. LYMAN.

The effect of intravenous and intraperitoneal injections of irradiated ergosterol. C. I. REED AND E. A. THACKER. *Am. J. Physiol.* 96, 21-7(1931).—Irradiated ergosterol, administered to dogs intravenously or intraperitoneally, is more effective than comparable doses by mouth. Toxicity and hypercalcemia are not parallel; hence the toxic factor may be something other than the vitamin D factor. There was no const. effect of the irradiated ergosterol dosage on the concn. of inorg. P of the blood. J. F. LYMAN.

The metabolism in pregnancy. V. The carbohydrate metabolism. A. W. ROWE, DOROTHY E. GALLIVAN AND HELEN MATTHEWS. *Am. J. Physiol.* 96, 94-100(1931).—*Glucosuria* is a common finding during the course of normal pregnancy. *Lactosuria* supervenes within a few days of delivery and may continue throughout the period of lactation or even beyond. Blood sugars are at low normal or slightly subnormal levels throughout pregnancy. The tolerance limit of 40 g. of galactose for the adult female declines to 20 g. during pregnancy, and sinks after delivery for 2 weeks or more to 10 g. With the cessation of lactation, or even before, the normal tolerance of 40 g. is regained. VI. The respiratory metabolism and acid elimination. *Ibid.* 101-11.—Pregnancy produces an increase in basal metabolism during the last 24 weeks of 13% more than that predicted from the change in wt. Following delivery there is a downward trend of significant proportions which reaches a min. value somewhere between the 3d and 5th week and then rapidly rises to normal. Alveolar CO₂ values indicate acid intoxication. Such a condition cannot be demonstrated with certainty, although there are minor evidences of possible acid retention. VII. The blood morphology. A. W. ROWE. *Ibid.* 112-4.—Hemoglobin shows a relative and abs. depression of moderate

degree during normal pregnancy. There is a moderate secondary anemia with lowered color index. J. F. LYMAN

Differential response of male and female ring doves to metabolism measurement at higher and lower temperatures. O. RYHOFF, GUINEVERE CHRISTMAN AND F. G. BENDER. *Am. J. Physiol.* 95, 111-20 (1930).—With an increase of external temp. from 20° to 30° the metabolism of male ring doves increased 20.1% and of females 23.3%. The differential response of the normal metabolism of the sexes to temp. is of consequence not only in metabolism studies, but in growth, vitamin and nutritional research. J. F. LYMAN

The relation between anorexia, anhydremia and gastric atony in dogs deprived of water. W. B. ROSE, C. J. STICKY, L. B. MENDEL AND G. R. COWELL. *Am. J. Physiol.* 96, 152-6 (1931).—Dogs fed an artificial ration complete in every respect were deprived of water for 6-day periods. A moderate increase in blood count with an increase in the hemoglobin value of less than 20% was associated with a definite decrease in the motor activity of the empty stomach and partial or complete loss of appetite. When atony of the stomach was recorded, there was a complete loss of appetite. Since anhydremia is a systemic disturbance other organs, particularly muscles, may fail to function with optimum efficiency in the presence of a moderate degree of anhydremia. J. F. LYMAN

Diet in relation to reproduction and lactation. III. F. D. WILKINSON AND V. E. NELSON. *Am. J. Physiol.* 96, 139-45 (1931).—Rats grew at a normal rate with 10, 20, 40 and 73.3% of cooked dried soy bean as the only source of vitamin B and C. Reproduction was normal on the lower levels of soy bean, but on the higher levels the rats were not so prolific. Satisfactory lactation was not obtained on any level of soy bean feeding unless a further supplement was added. When the soy bean made up 15% of the ration, liver from hogs and cattle and hog kidney, added to the diet, gave very marked improvement in lactation. Hog pancreas and beef pancreas, brain and heart gave some improvement while hog lung and spleen and beef thymus gave no appreciable improvement. The galactagog factor was destroyed in liver by drying for 12 hrs. in air at 150°. Ether ext. of liver possessed no galactagog properties. J. F. L.

Metabolic studies during pregnancy and menstruation. LEANE SANDFORD, THEODORA WHEELER AND W. W. BOOTHBY. *Am. J. Physiol.* 96, 191-202 (1931).—There is a gradual increase in the total heat production during pregnancy, amounting at term to 25%, and probably all accounted for by the added metabolism of the fetus, placenta and accessory structures. On an adequate protein intake of over 85 g. per day there is a pos. N balance varying between 0.5 and 1.1 g. daily. There is no evidence that the fetus grows at the expense of the mother's tissue if the diet is adequate. N retention results in a slight decrease in the % of total N eliminated in the urine as urea. There is also a slight increase in the ureic acid and amino acid N, an increase in creatine and no change in the creatinine eliminated until the ninth month when there is a sudden increase. Urinary NH₃ increases slightly. There are no significant changes in the more substances eliminated nor any changes in the chemistry of the blood. The mean value of 336 basal detus. of the heat production on a normal woman between 34 and 41 years of age is 33.1 ± 0.06 Cal per sq. m. J. F. LYMAN

Composition of a new diet for the study of B avitaminosis, a diet relatively poor in carbohydrate and rich in fat. MME. L. RANDON AND R. LECOC. *Compt. rend.* 192, 444-7 (1931).—Basing their work on previous expts. (C. A. 17, 1494; 18, 1322), R. and L. show the influence upon the development of polyneuritis in pigeons due to the feeding of a variety of carbohydrates and glucosides as a certain percentage of the diet. The compn. of the diet is: purified casein 8, purified filum 8, ovalbumin 8, butter fat 8, lard 18, carbohydrate 35, Osborne and Mendel salt mixt. 5, agar agar 8, filter paper 2. Comparative tests were made on diets differing only in the nature of the natural or purified carbohydrate, the following being used: glucose, levulose, maltose, saccharose, dextrin, corn starch, galactose, lactose and potato starch. All the pigeons submitted to this diet contg. glucose, levulose, maltose, saccharose or dextrin showed satisfactory temp., digestion and food absorption up to 20-30 days when acute polyneuritis and death occurred. On the diets contg. lactose, galactose and the starches, the onset of polyneuritis was delayed. N. M. NAVLOR

A direct quantitative relationship between vitamin A in corn and the number of genes for yellow pigmentation. P. C. MANGELSDOFF AND G. S. FRAYS. *Science* 73, 241-2 (1931).—Yellow corn is known to be much richer in vitamin A than white corn, and a similar assocn. between A and carotenoid pigments has been discovered in many other plant materials. Inheritance of seed color in corn follows definite Mendelian ratios resulting in the production of four classes of seeds, e. g., white, pale yellow, dilute

yellow and deep yellow, representing 0, 1, 2 or 3 genes for yellow pigment. Corn with these 4 genetic factors was bred and the seed was fed to rats in an appropriate diet. The following av. units of vitamin A per g. were found in seed grown in 2 different years when tested for A by the method of Sherman and Mansell (*C. A.* 19, 2517): 0 genes for yellow, 0.05 unit, 1 gene for yellow, 2.25 units, 2 genes for yellow, 5.00 units; 3 genes for yellow, 7.50 units. The results show (1) that a white-seeded variety of corn with little or no vitamin A in the endosperm is capable of forming A in its seeds if the genes for yellow pigmentation are introduced, (2) that there is a direct quantitative relationship between the no. of genes for pigmentation in the cells of the endosperm and the amt. of A in the seed. Each gene for yellow pigmentation causes the formation of approx. 2.5 units of A per g. of seed. The genes for yellow pigmentation are directly, or indirectly, responsible for the formation of A, with the production of carotenoid pigments as an intermediate step. This is the first case in which a direct relationship between different doses of the same gene and their chem. effect has been definitely established. It is improbable that this gene functions as an enzyme, as the total reaction resulting from enzyme activity is seldom related to the concn. of the enzyme. Each gene for yellow pigmentation, however, seems to govern the formation of a definite quantity of A. The straight line relationship between no. of genes and amts. of A is indicative of a direct chem. relationship between the genes and some other substance which is present in the endosperm of both white-seeded and yellow-seeded varieties of corn. C. H. RICHARDSON

Liver extracts as a source of vitamins B and G. W. D. SALMON AND N. B. GUERANT. *Science* 73, 243-4 (1931).—Early investigations of the distribution of vitamins showed that the livers of animals were a good source of vitamin B (vitamin G + B). The present expts. were made primarily to det. whether corn liver exts. afforded a good source of vitamin G for further concn. of the active factor. Growth studies on rats fed an appropriate diet in which various amts. of liver ext. were added indicated that the liver exts. are an excellent source of vitamin G, 0.10 g. per day furnishing sufficient of this factor for normal growth of a rat for a period of 9 weeks. Liver ext. is not such a good source of vitamin B. The ext. tested apparently contains 4-5 times as much G, but only about $\frac{1}{4}$ as much B as a sample of brewer's yeast. The vitamin B content of the ext. compared favorably with that of pure dry baker's yeast. C. H. RICHARDSON

Vitamins in sugar-cane juice and in some cane-juice products. E. M. NELSON AND D. BREESE JONES. *J. Agr. Research* 41, 749-59 (1930).—Sugar-cane juice is a poor source of vitamin B. The juice from the upper portions of cane stalks is richer in this vitamin than juice from the lower portions. Juice obtained from bagasse by using high pressure is richer in vitamin B than ordinary cane juice. Sugar-cane juice contains a small amt. of vitamin A but little, if any, vitamin D. Cane sirup, Louisiana and Porto Rico blackstrap molasses and cane cream, products made from sugar-cane juice, are devoid of demonstrable quantities of vitamin B. W. H. ROSS

The calcifying and toxic action of large doses of irradiated ergosterol in animals. Attempt to separate the two actions. H. SIMONNET AND G. TANRET. *Compt. rend.* 192, 586-8 (1931), cf. *C. A.* 25, 156).—The calcifying and toxic principles of irradiated ergosterol are resistant to oxidation and hydrogenation. Certain rabbits exhibit great ergosterol resistance, living several months, while other rabbits on the same dose die in a few days. Death resulted in 15-25 days on daily feeding of 20 mg. ergosterol irradiated 45 min. If irradiated ergosterol is administered in large doses 4 days a week, followed by 3 days rest, too rapid calcification is avoided, and a rabbit may live several months instead of a few weeks. The feeding 4 days a week of 0.2 g. of KI with 20 mg. ergosterol inhibited calcification and maintained a healthy condition in rabbits, while the omission of KI caused death in a rabbit in 223 days with intense calcification. Feeding 0.1-0.2 g. of irradiated ergosterol produced death in 6-8 days without appreciable calcification. This indicates that a toxic factor other than calcification exists and massive doses may kill the animal before any calcifying action appears. D. S. SEARLE

Chemical analysis and vitamin B determination in unpolished rice grains kept in tightly closed vessels (KONJO, OKAMURA) 12.

Handbuch der biologischen Arbeitsmethoden. Edited by EMIL ABDERHALDEN. Abt. 4, Tl. 14 (Lfg. 344). WILHELM PLÜCKER. Tabellen- und Rechenbuch für Nahrungsmittelchemiker. Berlin Urban & Schwarzenberg. 231 pp. M. 20, bound, M. 22.

MCLESTER, JAMES S. *Nutrition and Diet in Health and Disease.* 2nd ed. Philadelphia W. B. Saunders Co. 891 pp. \$3.50.

determining the rate at which a salt may penetrate the enamel. HENRY KLEIN *J Dental Research* 10, 727-31(1930), cf *C A* 24, 5343—The cond. changes in a tooth were followed by means of a cond. app. in which a vacuum tube amplifier was substituted for the usual telephone. The velocity with which a salt penetrated the enamel depended on the existing osmotic gradient, and was relatively slow in the canine tooth of the dog: thus 1.0 N KCl soln. required 8 hrs., and 0.1 N KCl soln. 24 hrs. J. S. H.

Basal metabolism. JOSEPH S. HEPBURN AND HARRY M. EBERHARD *Hahne-mannian Monthly* 65, 850-63(1930)—The theory and the technic of the basal metabolic rate detn. are considered, and an account is given of the clinical application of this test in gastrointestinal and related diseases. JOSEPH S. HEPBURN

Observations on certain physiological processes of the marmoset. II. Respiration. G. ENDRES AND H. TAYLOR *Proc Roy Soc (London)* B107, 231-40(1930)—CO₂ produced little increase in either the rate, the depth or the minute vol. of respiration of both hibernating and normal marmosets, unless the CO₂ content of the inspired air exceeded 5%. Higher concns. of CO₂ augmented the respiration, more markedly in normal (awake) marmosets than in those hibernating. III. Oxygen saturation of the arterial blood. G. ENDRES *Ibid* 241-5—In 2 hibernating marmosets, the O₂ satn. of arterial blood ranged between 93.5 and 96.4%. The temp. coeff. of the O₂ equil. of the erythrocytes was 1.8 on exposure to a partial pressure of 40 mm. CO₂, and 2.7 at a *pH* of approx. 7.76. IV. Blood sugar. *Ibid* 245-6—Blood was obtained by cardiac puncture. During hibernation the blood sugar of 5 marmosets ranged between 0.071 and 0.096%, during summer, its range in 3 marmosets was between 0.97 (a misprint for 0.097?) and 0.162%. JOSEPH S. HEPBURN

Isolation by cataphoresis of two different oxyhemoglobins from the blood of some animals. ALEXANDER GEIGER *Proc Roy Soc (London)* B107, 368-80(1931)—Two different oxyhemoglobins have been isolated from blood of the ox, sheep, ass, dog and man. They are best sep'd by electrodialysis followed by cataphoresis, and differ with respect to the O dissociation curve. JOSEPH S. HEPBURN

Relation between the anterior pituitary body and the gonads. III. Fractionation and dilution of ovary-stimulating extracts. MARGARET HILL AND A. S. PARKES *Proc Roy Soc (London)* B107, 455-63(1931), cf *C A* 24, 5819—Two types of reaction, (a) follicular luteinization and formation of atretic corpora lutea without ovulation, and (b) normal follicular maturation and ovulation, were produced in mice, rats, rabbits and ferrets by saline suspensions of pituitary tissue, alk. ext. of the anterior lobe of ox pituitary, and exts. prep'd from the urine of pregnancy. Extensive diln. expts with the urine exts. indicated that the 2 reactions could not be ascribed to the activity of a single hormone producing different responses at different concns. Some evidence of sep'n. of the luteinizing and the maturing activity was obtained by alc. fractionation of urine exts. Material sol. in 50% alc. was predominately luteinizing, that sol. in 30% alc. mainly produced follicular maturation. JOSEPH S. HEPBURN

Regulation of the ovarian cycle. MARGARET HILL *Science Progress* 25, 449-64(1931)—A comprehensive survey of the evidence that the hormone of the anterior pituitary body actually controls the ovarian cycle. JOSEPH S. HEPBURN

Influence of alteration in the vegetative system on the phosphorus metabolism. KÄTHE HESSE *Z. klin. Med.* 111, 729-41(1929)—In a series of metabolic balance expts., made in part on human beings and in part on dogs, it was shown that ergot amine and acetylcholine decreased both the P₂O₅ and the org. P content of the urine and both the inorg. P and the total acid sol. P content of the blood, while atropine and ephedrine had the reverse effect. No distinct influence was exerted on the Ca content of the urine, its vol. or its acidity. A relationship was thus found between the P metabolism and the functioning of the vegetative nervous system. J. S. H.

Biochemistry in relation to intelligence. H. D. POWERS *Science* 73, 316(1931)—Blood tests from a group of normal adults and one of idiots showed the Ca content within normal range in all cases. The phosphate content in the normal group ranged between 3.25 mg. and 8 mg. per 100 cc. blood, usually below 5.88 mg. and averaged 4.36 mg. The idiots had a high phosphate content, 5.98-12.45 mg., averaging 8.95 mg. per 100 cc. There was no reciprocal relation between the quantities of Ca and phosphate. MARY E. LEAR

The iodine content of animal organs. K. SCHARRER *Tierernähr.* 1, 563-77(1930)—A review. F. J. G. DE LEEUW

Problems of tropical physiology. W. BORCHARDT. *Ergeb. Physiol.* 31, 96-131(1931), cf *C A* 25, 1885—A review. C. M. McCAY

Physical-chemical equilibria of the blood. MAX HOCHREIN *Ergeb. Physiol.* 31, 421-506(1931), cf *C A* 24, 2779—A review. C. M. McCAY

Nutrition and metabolism of the cells of the eye. F. P. FISCHER. *Exptl. Physiol.* 31, 317 (1931) — A review C. M. McCAY

Recent developments in the field of iron metabolism. W. LINTZEL. *Exptl. Physiol.* 31, 344 (1931) et C. 4 24, 210, 3825 — A review C. M. McCAY

The inhibiting power of plasma from old animals on the growth of tissue culture. GINO BERGAMI AND G. FACCINI. *Riv. sci. ital. Biol. spec.* 5, 1075-8 (1937) — Under the exp'd. conditions used and within the limits of accuracy of the method, the inhibiting action of plasma from old animals on the growth of tissue culture *in vivo* has been confirmed. The transplanting of the testicle of a young animal to an old animal reduces the inhibiting action of the plasma from the old animal, without causing it to disappear in its entirety. When the testicular tissue has been completely absorbed, the inhibiting action of the plasma returns to normal. PETER MASCOCC

The presence of an estrus-producing substance in the male sexual gland. M. MAJNO AND B. FRATTINI. *Riv. sci. ital. Biol. spec.* 5, 1065-8 (1937) — Five kg. testicular glands was macerated in acidified water and pptd. repeatedly with salts of heavy metals at the isoelec. point. Five mg. of a cryst. substance was obtained which showed morphological characteristics identical with the pure follicular hormone. The needle-shaped crystals grouped themselves in the form of a tree, they were sol. in alc. and ether and somewhat less sol. in water. The cryst. substance in aq. soln. in doses corresponding to 250-500 g. of fresh gland was injected during a period of 3 days into 4 castrated rats. The animals were killed on the 5th day. The results were all pos.; all the preps. showed the 4 phases of the development of vesicular activity. Cross injections were then made. The activity of the testicular cryst. substance was tested on spayed female rats and the activity of cryst. folliculin on castrated male rats. All the female animals presented the typical estrus reaction, and all the male animals treated for 4 consecutive days with 5 units of folliculin gave typical pos. seminal vesicle tests. The hypothesis is advanced that the 2 hormones are either identical or have certain physiol. properties in common. PETER MASCOCC

The total sulfur content in human blood. STRATINO DEBANI. *Biochim. biophys. Acta* 18, 34-7 (1931) — In 20 individuals, the highest value of S in blood was 2314 g. per l. the lowest 1645 g. and the av. 1928. A. E. MEYER

The internal secretion of the thyroid gland and the development of the plumage in pigeons. I. Regeneration and moulting in hyperthyroidism. W. TH. LARIMONOV AND N. KOSOVA. *Biol. Zentr.* 51, 81-104 (1931) — In pigeons fed with thyroid substance, the regeneration of remoulted feathers takes place in a shorter time. Moulting could be induced. Changes in the thyroid gland of the birds were observed, presenting a transition to hypofunction. A. E. MEYER

Lehrbuch der allgemeinen Physiologie. Edited by ERNST GELLHORN. Leipzig: Georg Thieme. 741 pp. M. 47, half leather, M. 49.50.

PETER, AUGUST. Die Sekretionsmechanismen der Niere. Berlin: Walter de Gruyter & Co. 235 pp. M. 16 bound M. 17.50.

ZONDEK, BERNHARD. Die Hormone des Ovariums und des Hypophysenvorderlappens. Untersuchungen zur Biologie und Klinik d. weibl. Genitalfunktion. Mit einem Anhang: Die hormonale Schwangerschaftsreaktion aus d. Harn bei Mensch und Tier. Berlin: J. Springer. 343 pp. M. 38. linen, M. 49.60.

G-PATHOLOGY

H. GIBSON WELLS

The iodine content of the blood of patients suffering from cancer. F. S. FOW-WEATHER. *Brit. J. Exptl. Path.* 11, 400-7 (1930) — The claim has been made that the blood of patients suffering from cancer always contains notably less I than that of healthy persons. By use of the method of Leitch and Henderson (C. A. 21, 2339) the I content of the blood was det'd. in 12 normal individuals and in 24 cancer patients. In the normal cases the results obtained varied from 3.6 to 5.1 \times (\times = one-millionth of a g. per 100 cc.). Out of 24 cancer cases exam'd. only 11 show values below what may be considered as normal. This failure to find a const. assocn. between low blood I and cancer is definite evidence against the view that I deficiency is a cause of cancer. While there may perhaps be a tendency toward lowered blood I in cancer, there is at least as much justification for assuming that the reduction is the result of the cancerous condition as there is for assuming it to be the cause of such a condition. Further, since in rather less than half of the cases exam'd. was any reduction of blood I indicated, the detn. of blood I as a diagnostic test for the presence of malignant disease can have no value whatever. HARRIET F. ROAMES

A study of the blood pigment in obstructive jaundice, with observations on the van den Bergh reaction. W. J. GRIFFITHS AND G. KAYE. *Brit J Exptl Path* 11, 441-6 (1930). Cf. C. A. 25, 723.—Direct reaction human serum and the serum of rabbits after bile-duct ligation resemble in their properties a soln. in serum of the "direct-reaction" pigment shown to be present in bile. Human indirect reaction serum resembles in its properties a suspension in serum of pure bilirubin. The "direct reaction" pigment yields a diazo reaction over a wide range of pH, viz., when the pH of the mixt. after addn. of the reagent lies between 1.2 and 5.0 approx. on the other hand, the pigment of indirect reaction serum, like bilirubin, yields no prompt direct reaction until the final pH of the reaction mixt. is adjusted to about 6.0 by addn. of alkali to the serum before mixing with diazo reagent.

HARRIET F. HOLMES

Normal flocculating H- and granulating O-agglutinins for *Bacillus typhosus* in sera of different species. W. A. TIDDERMAN. *Brit J Exptl Path* 11, 447-55 (1930).—Normal agglutinins for *B. typhosus* in sera of man, guinea pig, rabbit, horse, sheep, goat and buffalo can be of the finely granulating O-type and the coarse flocculating H-type. It appears that in man and ruminants normal H-agglutinins occur more frequently than normal O-agglutinins, contrary to what has been observed in horses and rabbits. Sera of guinea pigs show the lowest content of normal typhoid agglutinins in the ascending scale sera of rabbits and man follow, reaching to horses and ruminants at the apex. After electroosmotic fractionation of the serum, both types of normal agglutinins for *B. typhosus* are found in the euglobulin fraction exclusively.

HARRIET F. HOLMES

The response of grain-fed pigeons to substances effective in pernicious anemia. J. M. VAUGHAN, G. L. MÜLLER AND L. ZETTEL. *Brit J Exptl Path* 11, 456-68 (1930).—The group of anemias characterized by megaloblastic hyperplasia must be regarded as due to the deficiency of some substance or substances essential for the normal function of the bone marrow. The bone marrow of the healthy grain fed pigeon is likewise characterized by the predominance of megaloblastic tissue. Such pigeons were used therefore as testing material for various substances capable of relieving conditions in man, assocd. with megaloblastic bone marrow. The oral or intravenous administration of liver prepn. effective in anemia produced a consistent pronounced response of the reticulocytes and 2 subfractions prepd. from it which had received clinical tests also caused a similar response. Beefsteak was as effective as potent liver prepn. in causing a rise of the reticulocytes. No reticulocyte response was obtained to a liver prepn. which was inert for pernicious anemia, a liver ext., in which the active principle had been destroyed, vitamin B1 and 2, bistamine, NaCl and casein. Substances capable of influencing the reticulocyte level are without effect on the total red blood-cell count. The grain-fed pigeon may perhaps suffer from a mild dietary deficiency, which is possibly a factor of importance in permitting its bone marrow to respond to certain substances.

HARRIET F. HOLMES

Diphtheria toxin-antitoxin floccules. F. C. SMITH AND J. MARRACK. *Brit J Exptl Path* 11, 494-502 (1930).—The expts. described give further support to the authors' contentions that diphtheria antitoxin consists of slightly modified pseudoglobulin mols., and forms the greater part of toxin-antitoxin floccules. There is no evidence that acid or base-binding groups are involved in the flocculation process, or that any appreciable amt. of heat is evolved. The amt. of protein in the floccules obtained from a balanced mixt. of diphtheria toxin and antitoxin is unaffected by the addn. of non-specific proteins, even when a special antitoxin soln., of which 1.3 was pptd. in the floccules, was used. Diphtheria antitoxin (horse serum) is pptd. by an antiserum to horse pseudoglobulin, the ppt. so formed combines with diphtheria toxin. No change of H-ion concn. occurred when purified and concd. toxin combined with a concd. and specially pure antitoxin soln. The isoelec. point of diphtheria toxin antitoxin floccules lies between pH 5.1 and 5.6. The heat of reaction of diphtheria toxin and antitoxin is not of the order found by Bayne-Jones (C. A. 19, 3109).

H. F. HOLMES

The concentration of the protective substance in antipoliomyelitis serum. W. T. J. MORGAN AND R. W. FAIRBROTHER. *Brit J Exptl Path* 11, 512-22 (1930).—From the exptl. results obtained with the immune serum from one horse immunized with living poliomyelitis virus, it would appear that the protein pptd. at a relatively low concn. of $(\text{NH}_4)_2\text{SO}_4$ is uniformly much more potent in antiviral action per unit weight of protein than that pptd. at higher concns. of this salt; and in conformity with this result the protein thrown out of soln. by dialysis of the immune serum, and the serum protein that is insol. in satd. NaCl soln., have been shown to possess potent antiviral properties. The evidence indicates that in this serum there was a general distribution of the antibody throughout the serum proteins, since carefully fractionated prepn.

as the pseudoglobulin and albumin possess definite although slight properties for neutralizing virus. HARRITT F. HOLMES

The appearance and development of certain antibodies in horses immunized with diphtheria anatoxin. A. REISEFELMAN, G. RABIN AND F. DELPOTTER. *Arch. intern. med. exp.* 5, 555-556 (1930).—A Schick pos. horse immunized with diphtheria anatoxin developed a max. antitoxin titer more rapidly but produced a smaller quantity of antitoxin than a Schick neg. horse. No agglutinins for the diphtheria bacillus were formed. The max. concn. of complement fixation antibodies was reached rapidly, but was maintained for only a short period. Complement fixation cannot be used as an index of the degree of immunity, since there appears to be no correlation between the production of antitoxin and fixation antibodies. F. R. MAIR

The role of serum globulins in the Wassermann reaction. S. T. WALTON. *J. Lab. Clin. Med.* 16, 451-6 (1931).—Normal human serum or solns. of serum globulins obtained by salt pptn. and dialysis appear to increase hemolysis in the complement titration of the Kolmer Wassermann procedure. The addn. of a small quantity of normal hum. in serum to the tubes employed in the titration of the complement allows for the affinity which the serum proteins may have for complement and makes possible a more delicate adjustment of the hemolytic system. F. R. MAIR

The influence of pH on the Sachs-Georgy reaction. S. SHIRAKAWA AND J. MIELKE. *Med. Dermatol. Soc. Japan* 13, 111-24 (1931).—See C. A. 25, 779.

T. H. RINEHART

GRAF 1. Die Krankheiten des Stoffwechsels und ihre Behandlung. Berlin: J. Springer, 599 pp. M. 20.60.
SCHMIDT WILHELM. Technik, Klinik und Theorie der Kolloidfreaktionen der Rückenmarksfüssigkeit. Dresden: T. Steinkopff, About 200 pp. About M. 15. bound, about M. 10. '30.

H—PHARMACOLOGY

A. S. RICHARDS

The physicochemical properties of tissues in relation to the normal state and the pathological state of the organism. VII. The action of certain amino acids on tar carcinoma of mice. F. VALLS AND A. DE COLORES. *Arch. phys. Biol.* 7, 181-201 (1930). Cf. C. A. 24, 2501-2. 587.—Various amino acids were injected into mice in which tar carcinoma had been produced. In the first series of exper. 13 mice received the following combinations of amino acids: aspartic acid, leucine, glycine, *dl* proline, *L*-cystine and *dl* alanine in equal quantities and in equimol quantities. Then a mixt. of *dl* proline and *L*-cystine in equimol quantities and one of equal quantities of *dl* proline, *L*-cystine and *dl* alanine were given and finally aspartic acid was used alone. In the great majority of cases an improvement was noticeable. As this series served only for orientation, 3 more series were studied, each one with 10 mice. These received equal quantities of *dl* proline and *L*-cystine and equimol quantities of *dl* alanine, *dl* proline and *L*-cystine. In no case did aggravation of the cancer occur, in a great number of cases improvement was noted. The improvement was strictly regression of the histologic type. Proline and cystine are undoubtedly of fundamental importance in this action. The favorable influence of *L*-cystine and *dl* proline as well as of *dl*-alanine is attributed to the fact that their isocn. points lie within a certain range. F. R. G.

Use of phenobarbital in infant feeding. ORVILLE BARBER. *Arch. Pediatrics* 44, 55-60 (1931).—Phenobarbital in doses of $\frac{1}{16}$ to $\frac{1}{8}$ grain at intervals of 4 or more hrs. has been administered to 76 infants between the ages of 1 week and 8 months for periods of 1 to 6 months. No cumulative or deleterious effects, mental or physical, resulted. Phenobarbital is safe and effective for the control of infantile vomiting, and more satisfactory than atropine. JOSEPH S. HERBERT

Biological estimation of some cardiotonic glucosides: ouabain, digitain, scillarenones and cyamaria. JEANNE LÉVY AND RAYMOND CAUEN. *Bull. in pharmacol.* 38, 23-37, 85-105 (1931).—The method of Hatcher Magnus on the dog was used for the comparison of samples of the cardiotonics. Three samples of ouabain proved to be equiv. if referred to the anhyd. substance or to crystals with $9\frac{1}{2}\text{H}_2\text{O}$. The min. lethal dose is given for the following compds: ouabain with $9\frac{1}{2}\text{H}_2\text{O}$ (anhyd. product shows $[\alpha]_D = -30.6^\circ$), 0.173 mg. per kg. animal, *g* strophanthus (20.07% H_2O ($[\alpha]_D = -28.48^\circ$), 0.1725, ouabain U. S. P. X with an av. of 11.1% H_2O ($[\alpha]_D = -30.6^\circ$), 0.1585 mg., equiv. to 0.178 of the crystd. compd. The international standard of strophanthin of 9110 has 0.173 mg. and the original contg. 10.94% H_2O has the value 0.1705. For tincture of strophanthus, used in a diln. of 1:100, it is between 0.04285 and 0.0047575 g.

strophanthus, the activity of 1 cc. corresponding to 3.50-3.97 mg. ouabain standard. The min. lethal dose of digitalin Natrvelle is 1.69, of digitoxin Merck 1.75, of digitaligenin 7.1 mg., of scillarene A (tested in 0.04% soln.) 0.353, scillarene B 0.167, com. scillarene C 0.288, tincture of squill 0.214 cc. (equiv. to 42.8 mg. of squill) and of cymarine 0.2302.

A. E. MEYER

Diphenyl ether series III Derivatives of the local anesthetic type (SUTER, OBERG) 10. *N*-Methyl *N*-phenylalkylaminoalkyl benzoates and *p*-aminobenzoates (COFF, McELVAIN) 10. Synthesis of some new thiazole amines containing the catechol group (OLIV, JOHNSON) 10. Synthesis of local anesthetics (KOVEX) 17. New derivatives of *p*-arsanilic acid (MORGAN, WALTOW) 10.

HANDOVSKY, HANS. *Pharmakologie, in ihrer modernen Problemstellungen*. Dresden: T. Steinkopff. About 200 pp. About M 14, bound, about M 15.50.

MOUTEVYRAT, J. *Contribution à l'étude du traitement de la syphilis par le bis-méthyl arsénic*. Étude de nouveaux dérivés. Paris: Le Français. 178 pp. F 25.

I—ZOOLOGY

R. A. GORTNER

A method for determining basal metabolism of fishes. MAIZE ADKINS. *Proc. Soc. Exptl. Biol. Med.* 28, 259-63 (1930).—A technique is described whereby fish can be kept at a fairly constant rate of metabolism while undergoing physiological experimentation.

C. V. B.

The role of bacteria in the nutrition of protozoa. J. MURRAY LUCK, GRACE SHEETS AND JOHN O. THOMAS. *Quart. Rev. Biol.* 6, 47-58 (1931).—A review with extensive bibliography.

J. B. BROWN

Evidence from the effects of potassium cyanide for linkage between polar growth, electric potentials and cell oxidation on *Obelia*. HILDA F. ROSENE AND E. J. LUND. *Pub. Puget Sound Biol. Sta.* 7, 337-44 (1930).—The growth of isolated internodes of the hydroid, *Obelia*, can be reversibly inhibited by cyanide. The effect is similar to the effect of HCN on cell oxidation and elec. potentials in the stem. An appropriate application of cyanide tends to reverse the polar sequence of growth of individual internodes. The effect of the KCN in a quant. way corresponds to the effect of an elec. current on growth. The continuous application of KCN will decrease the rate of growth, the magnitude of this depression being greater in young tissue. Additional evidence is presented for the quant. linkage between mechanisms of cell oxidations, bioelec. currents and the mechanisms of growth.

C. R. FELLERS

Investigations on the physicochemical properties of the hemocyanins of the octopus and of the horse-shoe crab. JEAN ROCHE. *Arch. phys. biol.* 7, 207-20 (1930).—No difference was found between the physicochem. properties of the natural hemocyanins and their decompos. products deprived of Cu by treatment with dil. acids. It is therefore not necessary to admit the presence of a special (prosthetic) group like the hema-tin in hemoglobin. The opinion of Hantzke that hemocyanin is a Cu protinate is probably nearer the truth. The isoelec. point of the hemocyanin of the octopus is pH 4.8 which is close to that of the crab, the lobster and the shrimp. The hemocyanin of the horse-shoe crab, however, has an isoelec. point of pH 6.2-6.4.

F. R. G.

Metabolic changes associated with pigmentary effector activity and pituitary removal in *Xenopus laevis*. I. Respiratory exchange. ENID CHARLES. *Proc. Roy. Soc. (London)* B107, 496-503 (1931).—The rate of respiratory exchange in the amphibian *Xenopus laevis* in air increases with a rise in the external temp., the O consumption per unit surface decreases as the body wt. increases. Females apparently have a lower respiratory rate than males. The respiratory rate is higher in unanesthetized eyeless animals than in eyed animals, and is lower in animals kept in total darkness prior to experimentation than in those kept in the light. Partial or total hypophysectomy considerably decreases the rate of respiratory exchange, this effect is more marked on the pulmonary rate than on the dermal rate. Hypophysectomy increases the R.Q. The rate of O absorption is from 2.5 to 3 times greater by way of the lungs than by way of the skin. II. Calcium and magnesium content of the serum. *Ibid.* 504-10.—The serum Ca and serum Mg were higher in females than in males and were not influenced definitely by pigmentary effector activity. Removal of either the anterior lobe or both lobes of the pituitary produced a persistent decrease in the Ca content of the serum.

JOSEPH S. HEPBURN

Osmotic changes in some marine animals. R. MARGARIA. *Proc. Roy. Soc. (London)* B107, 606-24 (1931).—In the crabs, *Maia squinado*, *Portunus depurator* and

(1929)), or from treatment for 48 hrs with 1% $HgCl_2$, 2% cresol, 5% phenol or 5% I suspensoid Merck. Acid or alkali was almost without effect. K. V. THIMANN

Change of vitality with age as based on the living unit of organism. I. Oxygen consumption in the daphnid, *Simcephalus caspinosus*. ARAYA TERAO *Proc Imp Acad (Tokyo)* 7, 23-5(1931)—By use of the data of Obreshkova (*C. A.* 25, 543), the following formula is obtained, $y = 65.855 e^{-0.0012x}$, where y represents the % rate of O consumption compared with the first instar young (size considered), x age in days and e base of Napierian logs. C. J. WEST

12—FOODS

F. C. BLANCK AND H. A. LEPPER

The reciprocal action between metal utensils and foodstuffs. II. A study of the biological value of metals. J. SCHWABOLD AND F. FISCHLER *Biochem. Z.* 232, 240-53(1931), cf *C. A.* 25, 1913—The method of this investigation consists in studying the effect on frog tadpoles of being reared in dishes made from different materials. No effect of any sort was observed with Al. Expts with Zn dishes show that under the influence of the natural corrosion products of Zn the growth and development of the tadpoles may be completely inhibited, however, if the action is not too far gone the condition is reversible. Expts with iron receptacles show that loading the tadpoles with corrosion products of this metal during the entire life span has had no deleterious effect and there is even some indication of a possible favorable action. The expts with tin utensils were very few but they suggest that this metal is either without action or its effect upon the organism is very slight, in consequence of its poor solubility and absorption into the organism. Copper, in these tadpole expts, has shown a surprisingly serious effect, even very minute quantities of its corrosion products being so toxic that the animals invariably died in a short time. S. MORCULIS

The use of aluminum for kitchen utensils. VICENTE COLOMBARO *Rev form (Buenos Aires)* 73, 75-87(1931)—The quantity of Al dissolved from an Al container of 1400 cc capacity, when 1000 cc of substance was boiled for 1 hr, was detd. In 100 g of substance the following values for Al (in mg) were found: distd H_2O none, tap water 0.04, 7% AcOH 15.66, 0.25% tartaric acid 6.11, 3% citric acid 6.47, 4% NaCl soln 2.35, 4% NaCl and 7% AcOH 13.15, milk 1.60, olive oil 5, sugar soln 2, compote 0.54, tomato preserves 45.6, broth 1.2, and white wine 1.83. The present knowledge of the toxicity of Al does not permit conclusions concerning the innocuousness of its use in the kitchen. A. E. MEYER

Cichorium intybus L. as food. I. FISCHNER *Kisérlet Közlemények* 31, 253-9 (1928)—Its use and composition are discussed. S. S. DE FINÁLY

Hungarian rye flours. Z. DE SÁNDOR *Kisérlet Közlemények* 31, 470-9(1928)—Ryes showed 13.62-14.21% water, 1.58-1.69% ash, 1.60-1.69% total N, 3.36-3.66 acidity no. and 9.12-9.63% albumins. Flours from these ryes contained 10.60-13.96% water, 0.44-3.18% ash and 0.78-2.85% total N and had acid no. 1.08-10.01. S. S. DE FINÁLY

Studies on improved Hungarian wheats. S. GÖMÖRY AND L. PAP *Mezőgazdasági Kutatók* 2, 116-39(1929)—The H_2O content was 9.21-10.81%; hl wt, 79.45-85.10 kg. No connection was found between hl wt, wt of 1000 grains and quality and quantity of gluten. The best sample contained 56.5% gluten. S. S. DE FINÁLY

Testing new wheat varieties. R. C. SHERWOOD *Cereal Chemistry* 8, 168-75 (1931)—Several new wheat varieties have appeared in recent years which have not been entirely acceptable to millers and chemists. Marquis is a variety of spring wheat that has met with great favor, though it is frequently grown under conditions which result in low flour yield, poor color, high ash and small loaf vol. Environmental conditions are liable to cause greater variations in water absorption than the differences between varieties grown under comparable conditions. Kota wheat yields well under dry conditions, but lodges badly and is susceptible to smut. Quality wheat is not equal to Marquis in baking properties. Garnet yields large loaves with only fair texture and color. Ceres is resistant to rust but susceptible to scab and smut. Reliance, Supreme and Renard are not resistant to stem rust. Minhard and Minturki have winter hardiness. Marquillo is a rust resistant variety, which seems to be better adapted to conditions in northern Minnesota than farther south. L. H. BAILEY

Apparatus to determine quality of wheat. K. KONORI *Mezőgazdasági Kutatók* 3, 426-31(1930)—Wheat is ground to flour, then stored for 8-10 days and worked

into a pulp which is examined in a machine on the principle of Rejto's pulp-examg. app. Numbers obtained should be compared with data for standard wheat types

S. S. DE FENLEY

Hygroscopicity of wheat. L. PAR *Mezőgazdasági Kutatók* 3, 349-54(1930).—200 Hungarian and 86 other wheat samples were examd. Two g. of sample placed on a surface of 25 sq. cm. showed 1.23% change in water content after remaining there for 4 weeks. The influence of hysteresis is excluded by producing a sample of 1.2-13% water content and keeping this hermetically sealed. The change in water content averaged 0.45-0.41%. Variations of hygroscopicity depend chiefly on climate. Drying or moistening has no influence on hygroscopicity

S. S. DE FENLEY

Wheat and flour studies. XVIII. A study of the nature of the acid responsible for the increase in acidity which occurs in flours during storage. ARNOLD H. JOHNSON AND JESSE GREEN *Cereal Chemistry* 8, 134-45(1931) cf *C. A.* 25, 1595.—The H-ion concns. of water exts. of stored flours extd. with ether were the same as those of the ext. of freshly milled flours. The idea is expressed that ether-extractable acids are alone responsible for the changes in H-ion concn. or in acidity (as measured by any method) which occur in flours during storage under proper conditions. The acids removed from flour by extn. with ether must, on the average, have rather long C chains, as their coeff. of distribution between water and ether was of the order 0.014 to 1. Only relatively insignificant quantities of lactic acid can be present in flour. L. H. B.

Natural chlorine content of wheat flours. L. PAR *Kis-érlet Közlemények* 31, 480-6(1928).—The Cl content was detd. by dissolving flour or ground cereals in 20% H_2SO_4 and pptg. Cl with AgNO_3 . The amount of Cl varied between 0.002 and 0.0738% in 32 Hungarian samples. Chlorinated flour can be identified by detg. the Cl content of a petrol-ether ext. The fat of untreated flours contained 2.5-3.5 parts Cl per million parts of flour, the fat of flours treated with Cl contained about 38% of the Cl used for the treatment.

S. S. DE FENLEY

A study of some methods of examining flour with special reference to the effects of heat. I. Effect of heat on flour proteins. C. W. HERR *Cereal Chemistry* 8, 1-23 (1931).—Kent Jones' viscosity curves in 10,000 auramine are confirmed, but it is shown that bacterial development and increase in acidity take place, and, therefore, no measurement is afforded of the proteolytic activity of the flour itself. Viscosity changes, however, may give a measurement of the resistance of the proteins to proteolytic attack. Incubation in pH 3.0 lactic acid did not completely inhibit bacterial growth, the viscosity curves are very different from those in water. Incubation in the presence of 1% NH_4F soln. gave apparently no bacterial growth, the viscosity decrease in the untreated flour was less and the viscosity increase in the strongly heated flour was also less. Viscosity acidity curves show interesting differences in the unheated and the strongly overheated flours, giving evidence of marked phys. changes in the flour proteins. Many of the well-known bactericidal substances were found to be insufficient for inhibiting the growth of bacteria in flour water suspensions. Various methods for the estn. of glutenin gave fairly good agreement among one another. The sum of the 3 fractions of proteins by recognized methods of analysis did not agree in many cases with the total protein figure detd. directly on the flour; in good-grade flours the sum is usually lower than the total, whereas in the lower grade flours the sum is frequently higher than the total. Different salts have different peptizing effects on some of the flour proteins, a subsequent alc. extn. will vary according to the preliminary salt treatment, indicating increased resistance to peptization. Ultra-violet radiations do not apparently fulfill any useful purpose in cereal analysis. If a large proportion of the natural moisture of a flour be removed at a low temp. and the heating then performed, the effects on the flour proteins will be produced but at a very much slower rate. II. Effects of heat on flour enzymes. *Ibid.* 145-61.—The diminution in phosphatase activity observed by Berliner and Rüter (*Z. ges. mikrobiol.* 4, 209-16(1928)) was not confirmed in flour heated for 10 hrs. at 180° F. or less. Proteoclastic activity of flour itself, as evidenced by amino N produced in an incubating ext., is of small value and can have little influence in a fermenting dough. The addn. of 0.2 g. yeast to 100 cc. ext. increased the proteoclastic activity. Heating may increase the resistance of proteins to enzymic attack. Catalase activity of flour suspensions is reduced by severe heating. Confirmation is obtained that the change in H-ion concn. of a dough during fermentation is of small magnitude. It is suggested that the pH of the aqueous phase, however, may alter and thus affect the enzymic activity. If a large proportion of the natural moisture of a flour is removed at a low temp. and then the heating performed, the enzymic activity appears to be inhibited at practically a normal rate.

L. H. BAILEY

A study of commercially milled flours dealing with protein and its relation to peptization and baking strength. R. H. HARRIS. *Cereal Chemistry* 8, 113-33(1931)—A series of 31 commercially milled flours of different baking strengths was baked and analyzed. The proteins of these flours were peptized by 0.5 N solns of $MgSO_4$ and KBr . Two methods of baking were used: the basic method, employing flour, sugar, salt, yeast and water, and a method which included 1% malt and 0.001% $KBrO_3$ for the weaker flours, and 3% malt and 0.5% Arkady for the stronger flours in addn to the basic formula. Total and nonpeptized protein gave very high relationships when correlated with loaf vol and would appear to be equally useful as a means of forecasting baking strength for this series. The percentage of total protein peptized was less significantly related to loaf vol than total or nonpeptized protein and would be of less value in predicting baking strength. Peptized protein did not appear to be of any great practical importance in its relation to baking strength. These conclusions are similar to those obtained in a previous study (cf *C A* 25, 1596). A series of 20 millstream flours was baked by the basic method and by a method which included 3% malt and 0.5% Arkady. These flours were analyzed and peptized by 0.5 N $MgSO_4$, KBr and KI . The results were considered in 2 groups: one comprising 12 millstreams which appeared to be of good baking quality and the other the whole group of 20 millstreams. The use of an improper method did not seem to be justified in evaluating the strength of millstream flours. Total protein is not a reliable index of baking strength when low grade millstreams are concerned. Nonpeptized protein is more significant in its relation to baking strength than total protein in millstream flours including the low-grade streams. The percentage of total protein peptized appears to be equal to, or better than, total protein or nonpeptized protein for purposes of predicting baking strength of millstream flours which include the lower grade streams. The "optimum coagulation" theory appears to be supported by the peptization data from the millstream flours.

L. H. BAILEY

Effect of moisture content of flour on heat of imbibition developed during the mixing of bread dough. LILLY GREWE. *Cereal Chemistry* 8, 162-5(1931)—The heat of imbibition of suspensions and bread doughs varies with the moisture content of the flour. Flours with moisture contents of 13.1 and 8.7% produced finished doughs which differed 33° in temp., a difference of sufficient magnitude to be important in bread making.

L. H. BAILEY

The variability of loaf volume in experimental baking. ALAN E. TRELOAR AND R. K. LARMOUR. *Cereal Chemistry* 8, 95-113(1931)—A statistical study of factors that may cause variations in loaf vols in exptl baking. "Molding personality" may be one cause of variation, and variation in molding technic may possibly be another cause. The absence of sugar from a dough of a standard flour resulted in loaves of approx 1/2 lower vol., but of good texture. The variability of the replicates was, however, doubled. The addn of 1 cc of 0.1% $KBrO_3$ soln to the mix in the absence of sugar increased the loaf vol., reduced the loaf wt. and reduced the variability of replicates. Range of temp. of the dough from 28° to 32° as it came from the mixer had no measurable effect upon the vol. of the loaves. At least 3 times the av. difference between duplicates should be taken as the min. difference to be regarded as significant between vols of single loaves baked from different flours.

L. H. BAILEY

The behavior of bread under the influence of some volatile substances. L. KARÁCSONYI. *Mezőgazdasági Kutatók* 3, 192-7(1930)—Aldehydes, ϵ , ϵ , $HCHO$, AcH , propionaldehyde, butyraldehyde, valeraldehyde, crotonaldehyde, BzH , salicylaldehyde, cinnamaldehyde and 2-furaldehyde, retarded staling of bread. The retarding action is proportional to the volatility and soly in H_2O . Some aldehydes change the consistency of the crumb, others color it. Proteins chiefly are colored, starch and probably fat remain unchanged. Aldehydes cannot be removed from proteins as easily as from starches.

S. S. DE FRIÁLV

Refrigeration in the bakery. A. R. FAYED. *Refrigerating Eng* 21, 251-5(1931)—The applications of refrigeration and air conditioning to the baking industry are explained. The process of bread making in a standard plant is described in detail.

A. H. JOHNSON

The quantitative measurement of carbon dioxide evolved in and lost from simplified muffin batters. ISABEL T. NOBLE AND EVELYN G. HALLIDAY. *Cereal Chemistry* 8, 165-7(1931)—Four types of baking powder were used in this study: tartrate, phosphate and 2 combination powders contg. different relative proportions of S. A. S. ($NaAl$ sulfate) and phosphate. When combined with flour, fat and water into a smooth batter, all the types of baking powder evolved less CO_2 than when combined with water only. The difference was greater in the case of the tartrate than with the other powders.

All types of powders also evolved slightly more gas when water was used in prep than the batter than when milk was used. The addn. of sugar up to 0.8% of the ingredients did not have any consistent effect upon the evolution of gas. Flour, fat, baking powder and water mixes lost between $\frac{1}{2}$ and $\frac{3}{4}$ as much CO_2 during their prep. as did the corresponding baking powder and water mix. When the tartrate and phosphate powders were used the batters lost more gas than when the combination powders were used.

L. H. BAILLY

Quality of Hungarian milks in 1928. I SZANYI *Kisrlet Közlemények* 32, 162-76 (1929)—Average values for 4148 samples of market milk from counties Moson, Vas and Zala were d 1.0310, dry matter 12.50% and fat 3.61%.

S. S. DE PINÁLY

Composition of Hungarian market milks. I SZANYI *Kisrlet Közlemények* 31, 309-20 (1928). The av. of 1301 samples of market milks from counties Moson, Vas and Zala was d 1.0313, dry matter content 12.5% and fat content 3.77%.

S. S. DE PINÁLY

Composition of milk of Hungarian Kammwoll sheep. J. CSIZSA *Kisrlet Közlemények* 31, 287-91 (1928). Dry matter averaged 20.11, fat 8.71, albumin 6.36, lactose 4.23 and ash 0.89%. The ratio of casein to albumin averaged 1.74. The milk is very similar to that of Merino sheep in compn.

S. S. DE PINÁLY

Composition of milk of Hungarian Allgauer cows. I SZANYI *Kisrlet Közlemények* 31, 241-51 (1928). The av. of 739 samples was sp gr 1.0336, dry matter content 12.7% and fat content 3.3-3.5%. *Ibid.* 32, 99-111 (1929).—The quantity of milk and its fat content show larger fluctuations than those of Summenthal cows. 31 to 26.3 l milk contg. 2.5 to 5% fat was obtained per day and cow. Morning milks contain less than 3.3% fat and therefore could only be sold after mixing with the evening milk.

S. S. DE PINÁLY

Some analyses of milk from Indian cows. TH. W. J. VAN MARLE *Pharm. Tijdschr. Nederland Indië* 7, 276-8 (1930).—The milk from 55 cows of different breeds, of 6 stables in the neighborhood of Batavia, had an av. sp. gr. of 1.0310, acid value 6.0, fat 4.0 (Gerber), milk sugar 4.51, Cl 97, albumin 3.7, f. p. 0.528, η_{sp} 1.3435.

D. R. KOOLHAAS

Seasonal changes of characteristics of milk fats of cows and of sheep. A. FALTYN AND G. DRINÁZKY *Kisrlet Közlemények* 33, 62-76 (1930).—In the milk fat of cows the Reichert Meissl no. varies between 22.56 and 31.71 with an av. of 26.01. Its max. is reached in winter and the min. in summer. The lauric acid no. of Kuhlmann-Grossfeld is 17.0-20.0. The xylene no. of Italte averages 18.27. The Polenske no. varies from 1.8 to 5.8, it is about 6 to 8 times as much as the R.M. no., depending on the season. Hungarian litters seem to contain more caprylic acid. The milk fats of sheep show Reichert Meissl no. 22.6-39.11, lauric acid no. 15.77-21.04, Polenske no. 2.2-6.0. Milk fats cannot be identified on the basis of these data.

S. S. DE PINÁLY

The effect of dilution on the titratable acidity of cow milk. H. H. SOMMER AND JULIA MENOS *J. Dairy Sci.* 14, 136-55 (1931).—Diln. of the milk with H_2O lowers the titratable acidity, diln. with an equal vol. of H_2O causes an av. decrease of 0.0211%, diln. with 9 vols. 0.0623%. In solns. contg. Ca and phosphates in amts. representative of milk, pptn. of $\text{Ca}_3(\text{PO}_4)_2$ will occur during titration with NaOH, increasing the titer. A similar pptn. must be expected in the titration of milk. The greater part of the diln. effect is attributed to a decrease in the amt. of $\text{Ca}_3(\text{PO}_4)_2$ pptd. during the titration. There are indications that a relationship of Ca to casein, similar to that of Ca to phosphates, is involved in the diln. effect. A lower pH was observed at the phenolphthalein end point in the titration of dild. milk. This is attributed to a "protein effect" on the indicator, while the rapid fading of the end point is due to the pptn. of $\text{Ca}_3(\text{PO}_4)_2$. The fact that the titratable acidity of the coned. milk is higher than calcd. from the acidity of the original milk and the ratio of casein, in spite of the loss of CO_2 , is attributed to the reverse of the same factors in the diln. effect. The neutralization of milk or cream with lime of known alk. does not lower the acidity to the calcd. point because the increased Ca ion concn. causes greater pptn. of $\text{Ca}_3(\text{PO}_4)_2$ in subsequent titration.

J. C. JERRINGS

A test for the detection of milk unstable to heat. GUY A. RAMSDALL, WM. T. JOHNSON, JR., AND FRED R. FRYAN *J. Dairy Sci.* 14, 93-106 (1931).—A rapid test, also applicable to factory routine, for the detection of milk unstable to heat has been developed from the results obtained in making comparisons between heat stability values of the coned. products and the pptg. actions of numerous acids and salts on the whole milk samples when heated to the temp. of boiling water. The most satisfactory pptg. agent was found to be KH_2PO_4 . The actual coagulation was brought about through

the combined action of the added acid salt and heat. Data are also given showing the relationship between heat stability and min. concn of phosphate soln necessary to produce initial coagulation when added to milk in the manner described in the test. Milk that coagulates in the test is invariably of low heat resistance, as are also composites of this class. In grading on the basis of the phosphate test the elimination of the phosphate positive milk often results in obtaining a milk of higher stability toward heat. The phosphate no. is defined and its relation to coagulation time shown. There is no apparent relationship between pH and phosphate no. Milks from single herds are reasonably const. in phosphate no. J. C. JERGENS.

Data on the bacteriology of nauseous and bitter-tasting milk. J. CSIZLER. *Kisérlet Közlemények* 33, 272-8 (1930).—The bitter taste is caused by *Bact. fluorescent*, the nauseous taste by *Micr. pituitosus* and *Bact. aerogenes*. The origin of such microorganisms is probably the wash water of bottles and machines. S. S. DE FINAŁA.

Casein plastics—milk in industry. RONNAR DOON. *The Times Trade and Eng. Suppl.* 23, No. 664, 9 (1931).—A review. E. M. SYMONES.

The effect of feeding menhaden fish oil on the secretion of milk and the composition of butter fat in the dairy cow. J. B. BROWN AND T. S. SUTTON. *J. Dairy Sci.* 14, 125-33 (1931).—The feeding of menhaden fish oil lowered milk production, the % and total butter fat. After resumption of control diet recovery to normal took place slowly. The characteristic highly unsatd. acids of the menhaden oil passed into the butter fat in small quantities. The analytical const. of the butter changed to those of a mixt. of butter with menhaden oil. Several specimens of normal butter have been shown to contain small quantities (0.3%) of a highly unsatd. acid similar to arachidonic acid. J. C. JERGENS.

Why fat tests vary. D. H. BAILEY. *Creamery and Milk Plant Monthly* 20, No. 2, 72 (1931).—The fat content of milk may vary because of breed of cattle, stage of lactation, health of the animals, time of milking, weather and season of freshening. Excitement of the cows at milking time may also cause sudden variations in the fat content of the milk. A. H. JOHNSON.

The feathering of evaporated milk in hot coffee. RANDALL WINTAKER. *J. Dairy Sci.* 14, 177-88 (1931).—The feathering of evapd. milk in hot coffee is promoted by (a) characteristics of the milk, especially the salt balance, (b) coffee which is coned, (c) prolonged extn. of the coffee grounds, (d) the use of small quantities of milk in the coffee, (e) the slow, rather than fast, addn. of milk. An empirical feathering test was developed. Examin. of 52 cans of commercial evapd. milk indicated that feathering in hot coffee is not a common defect of the milk now offered to the public. J. C. J.

Iron content of liquid and reconstituted dry milk: a comparison. ESTER B. HARDISTY. *Arch. Pediatrics* 48, 108-16 (1931).—A review of recent researches with a bibliography (12 references). JOSEPH S. HERNERN.

Manufacture, microflora and chemical composition of "tarhó." O. GRATZ. *Kisérlet Közlemények* 33, 27-34 (1930).—Tarhó is a kind of sour milk made in Hungary at 40-50°. Its microflora consists chiefly of *Thermobact. lactis* Orla Jensen and *Streptococcus thermophilus* Orla Jensen. It is somewhat similar to yoghurt, etc. Analyses give water 89.25, dry matter 11.75, fat 1.40, casein 3.29, albumin 0.14, albumose peptone 0.34, milk sugar 4.50, lactic acid 2.00, alc. 0.37 and ash 0.85%. S. S. DE FINAŁA.

Composition of creams from Miskolc and environment. I. PÁSZTOR. *Kisérlet Közlemények* 32, 317-26 (1929).—188 cream samples of Miskolc (Hungary) were analyzed. They averaged 18-20% fat and 6.8-7.6% dry matter content. S. S. DE FINAŁA.

Composition of Hungarian curds, butters and creams. J. KUPITS. *Kisérlet Közlemények* 32, 179-84 (1929).—Water content of curds of Dunántúl was 69-73% in 1927 and 68-73% in 1928, 56 and 35%, resp., of samples were adulterated. Butters contained water 15.8-29.23 and 39%, resp., 71% of samples were found adulterated with water. Creams contained fat 12.5-19.5 and 17.2%, resp., 23.7% were adulterated. S. S. DE FINAŁA.

Ice-cream standards. WM. C. GEAGLEY. *Ice Cream Trade J.* 27, No. 3, 34 (1931).—The possibility is considered of standardizing ice cream in regard to fat, total solids, sucrose and various flavors and colors. A. H. JOHNSON.

How temperatures should be used to insure quality ice cream. H. F. JUDKINS. *Ice Cream Trade J.* 27, No. 2, 31-2 (1931).—The manner in which temp. affects ice-cream mix and ice cream in relation to the quality of the finished product is discussed. Defects in quality caused mainly by improper temp. are coarseness, iciness and sandiness. A. H. JOHNSON.

The selection and use of flavor in making chocolate ice cream. W. H. MARTIN. *Ice Cream Trade J.* 27, No. 4, 39-41 (1931).—Several com. chocolate syrups were analyzed.

and found to vary in their content of moisture, fat, starch, ash and sugar. The variations in these constituents of the com chocolate syrups had considerable influence on the properties of ice-cream mix and ice cream. Ice creams prepd with cocoa or chocolate liquor were superior in flavor to those prepd. with the com chocolate syrups.

A. H. JOHNSON

What effect has pasteurization on the freezing of mix? E. L. REICKART. *Ice Cream Trade J* 27, No. 2, 35(1931).—Pasteurization time and temp. are shown to have considerable influence on the freezing and whipping properties of ice-cream mix.

A. H. JOHNSON

Recent developments in refrigerating equipment. R. S. WHEATON. *Creamery and Milk Plant Monthly* 20, No. 4, 71(1931).—A review is given of the application to the dairy industry of developments in refrigerating research and equipment.

A. H. JOHNSON

Data on the composition of Hungarian butters. I. SZANYI. *Kísérlet Kőzlemények* 31, 45-52(1928).—Twenty seven samples of Dunántul (Hungary) butter were analyzed in 1927. The water content averaged 14.4%, fat 83.6%, R. M. no. 26.26, I no. 40.8 sapon no. 216.2 and α -olein 43.2.

S. S. DE FINÁLY

Data on the estimation of butter pastries. Á. MILLIG AND I. KORPÁČY. *Kísérlet Kőzlemények* 33, 349-58(1930). The estn. of butter pastries on the basis of fat analysis is very difficult since the quantity of fat is very small and contamination by the fat of wheat flour and by the greasing oils of the distributing machines and baking plates may influence the results. Nineteen different pastries were baked with various quantities (1-4%) of different fats and fat mixes. The fat content on the basis of the dry material, α , I no., Reichert Meissl no. and Polenske no. were detd. in the product. The α s of butter, margarine, lard, coconut oil, table oil and wheat oil increased as a result of the baking. The I no. also increased except for table oil and wheat oil. The Reichert Meissl and Polenske nos. decreased in each case. Expts. were carried out to det. how much greasing oil can get into the pastries. 0.22% grease was taken up in case of very liberal greasing. This amt. in relation to about 1-2% butter content may influence very strongly the fat characteristics. 352 butter and margarine pastries were examd. by the Food Control Station of Budapest in 1929-30, and it is proposed that at least 5% butter should be prescribed for butter pastries. The Reichert-Meissl no. is the most important characteristic, but larger amts. of fat are needed for its detn. A new method is being worked out to make possible detns. for 0.5-1 g. fat.

S. S. DE FINÁLY

Bacillus amylobacter interferes with the manufacture of cheese from pasteurized milk. O. GRATZ. *Kísérlet Kőzlemények* 33, 260-8(1930).—The manuf. of Emmenthal cheese may be greatly interfered with by swelling if the milk is contaminated with anaerobic, gas producing, spore forming bacteria, since they and their spores are not killed by heating to 60° for 30 min. Contaminated milk cannot be purified by pasteurization.

S. S. DE FINÁLY

The manufacture of processed cheeses. I. Change of water content during melting. I. PASZTOR. *Kísérlet Kőzlemények* 32, 492-4(1929).—The water content of Emmenthal cheese increased 88%, that of Romadour cheese 63%, during melting. The av. water content of first-class Emmenthal processed cheese is 42-46%; that of Romadour cheese, 50-55%. II. Controlling the hardness of processed cheeses. *Ibid* 33, 268-71(1930).—The hardness is mainly influenced by the water content. Citrate, citric acid and Na_2HPO_4 play only insignificant roles. The hardness of processed cheese is greater than that of the original cheese if the water content is under 40%, it is about the same at 40-43% water content, and the product is soft cheese over 46% water content. III. Determination of p_H of processed cheese by means of a quinhydrone electrode. *Ibid* 307-16.—Diln. of a cheese soln. greatly increases the p_H value. Cheese, therefore, should be mixed with only 1-2 parts water. Filtering is not necessary. The potential is const. for only 5-20 min. after addn. of the quinhydrone. the detn. should be made at once.

S. S. DE FINÁLY

Why stanniol paper turns black from processed cheeses. I. PASZTOR. *Kísérlet Kőzlemények* 33, 316-24(1930).—Expts. made with *Bac. amylobacter*, *B. putreficus* and with H_2S water proved that H_2S did not cause any blackening. Investigations showed that the blackening has 2 phases: (1) Sn is dissolved under the influence of the chemicals used to produce processed cheese, e. g., citrates. Soln. is helped by air and higher temps.; (2) Sn-Sb galvanic elements are formed and SnO produced by anodic oxidation causes the black color. This process seems to take place at a higher p_H , mostly above 5.8.

S. S. DE FINÁLY

The specific gravity and water content of coagulated cheese. The object of subsequent stirring in the manufacture of Emmenthal cheeses. K. VAS. *Kísérlet*

Kozlmenyek 33, 377-85(1930)—The detn of sp gr in salt soln is interfered with by diffusion. A mixt of CHCl_3 and $\text{C}_{12}\text{H}_{22}\text{O}_6$ evaporates quickly. A mixt of paraffin oil and CCl_4 is very viscous and retains air bubbles. Rapid sp gr detns can be made with the hydrometer scales of Nicholson. Investigations proved that the sp gr increased (from 1.074 to 1.073) and the water content decreased (from 58.84 to 40.13%) during subsequent heating and stirring. The object of the subsequent stirring is not to dry the cheese, but to bring about such a phys. state that the excess whey could be filtered off the cheese particles. The measurement of sp gr is therefore not enough for the evaluation of coagulated cheese. S. S. DE FINALLY

Analytical microscopy of commercial egg albumin. S. MENDELSON. *Chemist-Analyst* 20, No. 3, 4-5(1931)—Desiccated, com egg white may contain dried milk, casein, starch H_2BO_3 or $\text{Na}_2\text{B}_4\text{O}_7$. The turmeric test will detect borate and the I test starch. I examn under the microscope will show, first, the amorphous particles of albumin as translucent, irregular masses bounded by rounded edges. A drop of clove oil will obliterate these outlines. Red, yellow, black or brownish particles will reveal the presence of coagulated residues. By heating over a low flame the albumin will be fixed and colored yellow by the clove oil. In the absence of coagulated albumin, any dried milk particles will be observed in bold relief in contrast to a nearly clear field. Comparison under the microscope with samples of known purity is advisable. W. T. II

Chemical analysis and vitamin B determination in unpolished rice grains kept in tightly closed vessels. M. KONDO AND T. OKAMURA. *J. Sci. Agr. Soc. (Japan)* 1929, 183-204 (in German 205), cf. *C. A.* 23, 3493.—Rice kept 4 years in straw sacks showed marked decrease in the content of ash fats and proteins, whereas rice kept 4 years in sealed tin cans filled with either CO_2 or air did not show any decrease in those constituents. The vitamin B content of the rice kept for 4 years in sealed tin cans was the same as that of rice half a year old. Unpolished rice kept in sealed tin cans is free from insect attack and keeps its germinating power with little change in its phys. and chem. nature even after several years. K. SHIMONARA

The use of benzoic acid in chopped meat. TJADEN. *Arch. Hyg.* 104, 184-96 (1930)—Preps contg benzoic acid or Na benzoate can be used only with fresh meats since the preps do not remove evidences of decay nor inhibit bacterial growth already at its height. The addn of the preps to fresh meat tends to prevent the development of bacteria and appears to be a harmless practice. F. R. MANN

Composition of Hungarian honeys in 1928. J. SITZ. *Kutlatt Kozlmenyek* 32, 530-45(1929)—Thirty two samples from counties Baranya and Somogy showed an av. water content of 14-22.5%. Glucose and fructose contents were also detd. S. S. DE FINALLY

Albumin content of Hungarian honey. J. SITZ. *Méregszedési Kutatások* 3, 25-8(1930)—Peptones and globulin were present and some albumin. Protamines, ale-sol albumins, histones, albumoses and albuminoids could not be detected. Diastase, invertase and catalase were present. There were traces of A, B and C vitamins. S. S. DE FINALLY

Studies of methods used to detect heated honeys. G. H. VANSSELL AND S. B. FREEBORN. *J. Econ. Entomol.* 23, 428-41(1930)—Heat produces chem changes in extd honey. Overheating renders the honey undesirable and it is often condemned by European importers. Methods used to detect heated honeys are based upon (1) destruction of diastatic activity and (2) the production of furfural derivs. from levulose by heating. If diastatic activity is absent and furfural derivs. are present in small amts., the honey is declared to be overheated. The diastase test is of doubtful value (*C. A.* 24, 898) because the diastase of honey is not a product of the metabolism of the bee but is derived from flower pollen. Some honeys are so lacking in pollen that they are deficient in diastase and would be rated as overheated. The Fiehe test as used in Germany for hydroxymethylfurfural is made as follows: grind 5 g honey and dissolve in anhyd. H_2O . Evap. the Et_2O ext. at room temp and moisten the residue with a freshly prepd soln of HCl (d 1.19), 100 g contg 1 g resorcinol. A strong cherry red color lasting at least 1 hr indicates the presence of artificial invert sugar, a weak, quickly disappearing orange to pink color is considered to indicate overheating. Lampitt, Hughes and Rooke (*C. A.* 24, 1162) find that the Fiehe test is not positive until sufficient heat has been applied to the honey to change its color and flavor. V and F obtain the same results with freshly extd honeys, but extd honeys 1-3 yrs old that have been allowed to crystallize give a positive Fiehe test (German method) before heat has been applied. When honey crystallizes the levulose concn of material in soln increases 40-60%, it is assumed that the acid content also increases. This

greater concn. of levulose may account for the increased formation of furfural. The increase in time to give the reaction depends somewhat upon the levulose dextrose ratio. It may be hastened by increasing the acid concn. even when the honey has not crystallized. Conclusion: The diastase and Liehe tests are not reliable tests for over-heating unless the pollen count is taken into consideration. If the pollen count is 3000 pollen grains or more per g. honey and the diastase test is negative, the honey has been overheated. C. H. RICHARDSON

Further observations on the deterioration and spoilage of honey in storage. GEORGE I. MARVIN. *J. Econ. Entomol.* 23, 431 (1930).—Fermentation of honey caused by sugar-tolerant yeasts is one cause of its deterioration. Granulation of honey offers conditions conducive to the growth of the yeasts. The by-products of fermentations which cause off-flavor are CO_2 , alc. (which rarely exceeds 5%) and a non-volatile acid. Honey which is to be sold in small containers should be heated to 71° and sealed while hot, then cooled quickly. It should be stored below 12°. C. H. RICHARDSON

Freezing and melting points of fruits and vegetables. AURELY L. SMITH. *Refrigerating Eng.* 21, 272-3 (1931). The f. p's of the juices expressed from 10 fruits and 14 vegetables are given. The lowest f. p. found was 25.40° for grape juice and the highest f. p. was 31.82° for lettuce juice. A. H. JOHNSON

Abnormalities in the composition of oranges. P. R. V. D. R. COREMAN. *South African J. Sci.* 27, 310-6 (1930). cf. *C. A.* 24, 2208. To det. the compn. of a "normal" orange, analysis was made each week for 3 months of 24-30 oranges selected from different trees in the same orchard. Detns. were made for sol. solid, sugar and acid. Detns. of d. were made and converted to sol. solid. (*Cf. C. A.* 19, 1881; 23, 3487.) As the fruit develops, sol. solid and sugar content increase, while acidity decreases. It was thought that oranges from different trees would not show large deviation from the mean values, but the results indicated a wide variation in the compn. and hence the difficulty of obtaining a fruit of normal compn. The oranges from 2 of the trees gave very different analyses from those of the other trees with respect to (1) rate of change of sol. solids with time, and (2) av. values for sol. solids, sugar and acidity. By using the method of analysis of variance by Fisher, it was shown that the data obtained on oranges from these 2 trees should be omitted because of the deviation from the mean values. The results show the importance of analysis of the oranges in order to eliminate trees bearing abnormal fruit and thus insure the uniformity of the crop. N. M. N.

Hungarian condensed tomato juices. I. H. INV. *Kisérlet Közlemények* 33, 82-8 (1930).—Samples in 1928 contained 14-31% dry matter, in 1929 15-35% dry matter, with an av. of 24-25%, 56-61% of this consisted of invert sugar and 4.6-5.9% of oxalic acid. Hungarian juices are not salted or oiled artificially. S. S. DE FINÁLY

The analysis of ground coffee, in connection with the requirement of the Dutch food law. W. F. DONATI. *Pharm. Tijdschr. Nederland-Indië* 7, 310-2 (1930).—Several samples of ground coffee with good chem. const. were shown by microscopic examn. after boiling with HNO_3 and then with KOH to be roasted Indian corn. The most adulterated coffee had the highest ext., but the requirement of a max. ext. value does not insure purity. Besides chem. analysis, a microscopic examn. should be required, for this the app. of Eredam is recommended. D. R. KOCHMAAS

Bleaching almonds. A. A. RAMSAY AND G. W. NORRIS. *Agr. Gas. N. S. Wales* 42, 153-4 (1931).—Almonds off in color, but not stained black, were successfully bleached by immersing the nuts in 1% H_2SO_4 for 2-5 min., or in a 0.5% soln. for 5-10 min., followed by drying the nuts in thin layers. Almonds badly stained (black at the base of the nut) were successfully treated by the following methods: (A) The nuts were immersed for 20 min. in chloride of lime (1 oz. per gallon) acidified with AcOH , and after drying for 1 day were dipped in 1% H_2SO_4 for 5 min. and again dried. (B) Similar treatment to (A), but after drying the 2nd time the nuts were washed in several changes of water and dried as before. (C) The nuts were immersed for 5 min. in 1% H_2SO_4 , dried for 1 day and spread out in thin layers. They were washed with water and again dried, then dipped for 20 min. in chloride of lime acidified with AcOH and dried as before. K. D. JACOB

Determination of essential oils in spices according to Griebel. L. BENEDEK. *Kisérlet Közlemények* 33, 99-102 (1930).—A modification of the method of Griebel (*C. A.* 20, 3516). A 500-cc. shaking cylinder with glass stopper is used instead of a separatory funnel. Data agree to 0.1%. S. S. DE FINÁLY

Experiences on the quality of powdered pepper. I. HORVÁTH. *Kisérlet Közlemények* 32, 546-54 (1929).—Contents of ash, sand, sandless ash, crude fibers and CaCO_3 , and microscopic examn. were used as the basis of evaluation. A modification of the

method permits extn with alc in 3 hrs instead of 30 hrs. The ash content varied from 4.60 to 8.96%, sand 0.14-2.88%, alc ext 7.10-10.70%. S S DE FINÁLY

Capsaicin reaction and evaluation of paprika. K. FODOR *Kisfűrt Késkemén* 33, 155-78(1930)—With VOCl_3 capsaicin gives a blue color in solns of acetone, of ether or of chloroform and gives a green color in dil solns. This reaction can be used for the detn of capsaicin in paprika. Shake 2 g dry paprika with 10 cc acetone, then centrifuge and filter. Treat 5 g of filtrate with 0.4% coned HCl and 0.1 g NH_4O_2 . The color must be detd within 20-30 sec, as troublesome reactions take place. Another method is an ether extn made according to the method of RIGÉNYI and the extn is dropped in a soln of VOCl_3 . Evaluation of paprika should be made on the following basis: (1) color and taste, (2) microscopic examn, (3) content of capsaicin, (4) content of sand and ash and (5) ether extn. S S DE FINÁLY

Rapid determination of ether extract of paprika. I. HORVÁTH *Kisfűrt Késkemén* 33, 103-5(1930)—A modification of the method of D. MÖTÖR (*C. A.* 23, 4979). Shaking is done with a machine for 1 min and sedimentation is hastened by centrifuging for 15 min. The method is quite as reliable as that of RIGÉNYI, which is now in use. S S DE FINÁLY

Composition and use of "Old Hickory Smoked Salt." O. VARGA *Kisfűrt Késkemén* 33, 197-200(1930)—The product of a factory at Cincinnati, Ohio, was examd. It consisted of water 0.1, NaCl 94.5, Na_2SO_4 0.4, sand 0.1 and org. matter 0.9%. The taste of the product made with this salt was not so good as that made by actual smoking. The drying of the product required 4 days, during which microorganisms multiplied. Cocci and yeasts belonging to the genus *Torula* were found, further rod-like bacteria were in the flesh near the bones. Expts made in a large smokehouse showed that such salt mixes cannot quite substitute for actual smoking. S S DE FINÁLY

Possible uses for plant lecithin. G. HEINTZ *Z. Ernähr.* 1, 53-61(1931)—Com. lecithin obtained by the Bollmann method (*C. A.* 17, 3234) from soy beans consists of 60 lecithin and 40% fat. The plant product has the same chem. and phys. properties as that from eggs and is cheaper. It is possible to use this soy lecithin in place of egg yolk in baking. Five tablespoonfuls of a 20% soln correspond to 1 egg. F. P. G.

Composition and forage value of Sudan grass hay. A. ZAITSCHEK and L. DÖRNER *Kisfűrt Késkemén* 31, 336-9(1928)—The compn., digestibility and starch values of Sudan grass hay were not less than those of other Hungarian grasses. The digestible albumin content was as high as that of first-class hays. S S DE FINÁLY

Quality of alfalfa hay in relation to curing practice. T. A. KRESSELBACH and ARTHUR ANDERSON *U. S. Dept. Agr., Tech. Bull.* 235, 25 pp(1931)—Approx. 75% of the protein of alfalfa is in the leaves, hence N conservation in the curing process is concerned principally in leaf conservation. Field expts in Neb. showed that the more rapidly the hay can be cured, the better will be the quality. The internal moisture of alfalfa stems is lost by direct evapn. through the stem epidermis rather than by bleeding at the cut end or by continued transpiration after cutting. Turning or scattering the windrows of semi-cured alfalfa had little or no effect on color, leafiness or protein content. In a poultry feeding comparison, field-cured and artificially dried alfalfa hays showed no significant differences in vitamin A content. C. R. FELLERS

The conservation of fodder. ALLOIS KROCIK *Listy Cukrovár* 49, 290-4(1931)—Beet greens were ensiled and deposited in ground pits. The tightest packing did not exclude the presence of air, which leads to the formation of AcOH . Increase in temp. increased the butyric acid content. Four regions in the country were selected for the storage, only one expt. was highly successful. The beet greens in this instance were deposited in pits very rapidly and without interruptions so that the exposure to air was very short. Beet greens stored in concrete silos showed 1.7% lactic acid, beet greens from ground pits, 0.7%. Very small quantities of butyric acid formed in silage from concrete silos. In cases where the beet greens were exposed to air during a prolonged harvest, the AcOH increased to 1.8% in concrete silos. FRANK MARSH

The use of denatured sugar, molasses, molasses-fodder and sugared cossettes in agriculture. K. ŠANDERA *Listy Cukrovár* 49, 302(1931)—The feeding expts at Rostock, Germany, are reviewed briefly. FRANK MARSH

Mahua waste as food for milch cattle. K. HABIB HASAN and S. R. BHATE, *Proc. 15th Indian Sci. Congr.* 1928, 48-9—Mahua flowers after fermentation and distn. are fed to milch cattle in India. Feeding expts. to det. their food value, their effect on the health of the animals and the quality of milk gave on the whole favorable results. E. J. C.

Manufacture and quality of Hungarian fish meal. K. LUKÁCS. *Természetiud Közöny* 63, 115-20(1931)—Fish caught in Lake Balaton are dried in a vacuum at about 50° and ground to a meal contg about 8% moisture, 6-8% fat, 3% calcium phosphate and 62 to 68% raw albumin. The meal can be used as forage without extg. the fat. S S DE FIVÁLY

Forage value of washed and dried beets with leaves. V KURELEC. *Kisérlet. Közlemények* 33, 229-33(1930)—Digestibility of forage is greatly increased by a previous washing and drying of beets. Data of metabolism are given on the basis of expts with 2 sheep. S S DE FIVÁLY

An interpretation of the feeding standards for growing dairy cattle. J. B. FITCH AND R. H. LUSH. *J Dairy Sci* 14, 116-24(1931)—The Morrison standard for growing dairy animals can be readily used in the majority of exptl trials. A table is presented that enables this standard to be interpreted more quickly. J C JURRENS

Microflora of processed cheese (Csészár) 11C. New applications of low-temperature freezing (Zurbro) 13. The evaluation of the roots of fodder beets for eugenic purposes (Kóchar, et al) 11D. Al content of plants, chiefly food plants (Bertrand, Lévy) 11D. Apparatus for grading tea (Brit pat 339,758) 1. Spark-discharge apparatus for generating a bleaching gas for flour (Ger pat 521,360) 4.

Handbuch der Milchwirtschaft. Band II, Teil 1. Edited by W WINKLER. Berlin J Springer 488 pp M 48

"Cereal coffee." H OHLHAVER. Brit. 339,821, Jan. 2, 1930. The cereal used is caused to undergo 'self fermentation' during soaking, and is then roasted. Wheat may be soaked for 20 hrs in cold water or for 8-10 hrs in water at a temp of 25°

Pipe-system heater for preheating milk before filtering, etc. RALPH B. BAGBY. U S 1,797,312, March 24. Structural features

Apparatus for treatment of fruit with volatile agents such as sulfur dioxide. FRDD-ERIC A. EUSTIS (to Virginia Fruit Fumigating Co) U. S. 1,797,680, March 24. Structural features

Apparatus for treating food products (as in sterilizing milk) by direct contact with steam. GEORGE GRINDROD (to Grindrod Process Corp.) U. S. 1,797,769, March 24. Structural features of app with steam jets for direct injection of steam into the material treated. U S 1,798,120 relates to a similar treatment for sterilizing various food products, some of which, such as ice cream mixes and gelatin solns, may be coned after sterilization to a hydrated jelly like condition. App is described. Cf. C A. 24, 4101.

Treating citrus fruits to prevent stem-end rot and blue-mold rot, etc. HARRY R. FULTON and JOHN J. BOWMAN. U S 1,797,572, March 24. An aq soln. of Na aluminate (suitably of about 3% strength) is used at a temp. such as that commonly used for washing fruit, and the fruit is allowed to remain in contact with the soln for 1-10 min and then sprayed to remove most or all of the adhering soln. U. S. 1,797,573 relates to the similar use of a Aln sulfate soln.

Margarine, etc. EINAR LANGEFELDT and INTERNATIONALE GRADIN A.-G. Fr. 696,506, May 20, 1930. In processes for the manuf. of margarine, artificial creams or other emulsions rich in fats, the milk or other aq soln. is submitted to a pasteurization at temps between 50° and 90°, preferably about 65°, in the presence of CO₂, the liquid being heated to pasteurization temp after absorption of the CO₂ or the liquid heated to pasteurization temp is submitted to the action of gaseous CO₂ by bubbling, by agitation or by contact with thin layers of the liquid.

Smoking meat and fish. THE ATLANTIC COAST FISHERIES CO. Ger. 520,193, Mar 28, 1929. See Brit 309,405 (C A 24, 668).

Analytical apparatus for washing out starch from flour. "MAG" MÖHLENBAU UND INDUSTRIE A.-G. Ger 519,370, Oct. 18, 1929.

Dough ingredients for bread, etc. A. D. BLANK, H. A. KOHMAN and A. SCHULTZ (to Standard Brands, Inc) Brit. 340,072, Oct. 22, 1928. A prepn. suitable for use in making leavened products without making a sponge comprises a material of acid reaction such as H₃PO₄ or Ca acid phosphate, an NH₄ salt such as the chloride or phosphate, an oxidizing salt such as KBrO₃ or KIO₄, an enzymatic agent such as papain, pepsin, Taka diastase or a diastatically active malt ext. and soda for neutralizing acidity during baking. These ingredients may be mixed with a diluent such as flour to facilitate uniform distribution with the dough materials.

Flavoring agent for confectionery. DEUTSCHE GOLD UND SILBER-SCHNEIDEN STAHL WERK ROESSLER Fr 696,872, June 10 1930 A flavoring material, particularly for confectionery is made by the reaction of salts of glutamic acid, e g, Na glutamate, with appropriate org acids such as HCOOH , if necessary in the presence of solvents

Drying fruits, etc. BRANCH & SPEHR Fr 696,910, April 15, 1930 The vitamins are preserved in dried fruits, vegetables, etc., by treating the fruit, etc., with acids and salts before drying e g with H_2SO_4 and NaCl

Confectionery cream. WILLY BRITSCHOT Swiss 143,006, Feb 20, 1930 A stable confectionery cream consists of a hot concd soln of sugar, grape sugar, eggs and vegetable fat Harmless aromatic and coloring agents may be added

Flue-gas heated roller driers for the manufacture of potato flakes. BERNHARD SCHMIDT Ger 517,717, May 5, 1928

Radioactivated foods. GEORG SENFTNER Fr 695,729, May 16, 1930 Sugar crystal from radioactive water is added to the usual ingredients and finished in the usual manner

Cheese. JEAN M DAGAND Fr 695,760, May 16, 1930 Cheese is made by curdling milk, breaking it up into a large no. of particles, stirring, heating and afterward placing it in a mold to drain and finally pressing

Table salt. JEAN A ALLÈGRE Fr 695,563, May 14, 1930 A mixt of NaCl and Mg salts suitable for table use is made by adding a water absorbing substance which tends to adhere to the Mg salt and render it deliquescent, whereby the mixt does not become damp on exposure to the air. MgSO_4 is preferably used, but Na_2SO_4 may be used

Testing milk. N GERBER CO M. S. H. Ger 518,978, Aug 27, 1929 Adds to 516,423 (C A 25, 1922) The method of Ger 516,423 is modified by adding a peroxide to the indicator employed The test described in the prior patent is thus combined with the detection of colostrum through the evolution of gas due to its content of catalase

13—GENERAL INDUSTRIAL CHEMISTRY

HARLAN S. MINER

Research in industry—what it is and what it is not—application of knowledge of scientific principles to commercial production. P DUNNEATH *Electrician* 106, 543-5 (1931) —A discussion emphasizing the need of responsibility and of definite conclusions

What's patentable in industrial and chemical engineering? CHARLES W RIVISE *Ind Eng Chem* 23, 589-8 (1931) W H B G G

Cheaper power for the chemical industry. W S JOHNSON *Ind Eng Chem* 23, 474-5 (1931) —A description of the facilities for the generation of steam and electricity possessed by 2 modern plants B A SOTLE

Electricity and the plastic industry insulating materials. WM D OXEN *The Times Trade and Eng Suppl* 28, No 64, 22-3 (1931) —Examples, with illustrations, are given. E M SYMMES

The Bakelite Corporation. L V REDMAN and A V IL MOREY *Ind Eng Chem* 23, 525-7 (1931) G G

Water as absorption material. L PLATT *Kunststoffe* 20, 2-5 (1930) —A discussion of absorption by water of various materials including alc., cresol and acetone both in liquid and vapor phases, with vapor pressure and boiling curves of binary mixts. e g, EtOH H_2O and acetone water The literature is reviewed B H

Evaporation of water by hot dry air. R H NEWTON and T C LLOYD *Ind Eng Chem* 23, 530-2 (1931) —Graphs show the relative importance of water flow, air flow and air temp Water flow was found to be the most important variable L A PRIDGEON

Drying by spraying in the chemical industries. W KCHLES *Chem Rundschau Mitteleuropa u Balkan* 6, No 14, 100-1 (1929) —Liquid is sprayed and dried very quickly with warm air The solvent evaps instantly The app is described. S S DE F

Chemical industry of Bulgaria. E PETZALL *Chem Rundschau Mitteleuropa u Balkan* 6, No 14, 91-3 (1929) —General review with numerical and statistical data S S DE FINALLY

A general algebraic and graphical calculation of the technical work-processes with

mixtures of two substances. A. WEISE. *Z ges Kälte-Ind* 38, 17-20, 33-37 (1931) — The method is applied to the absorption refrigeration process. F. D. ROSSINI

Study of methods of determining moisture in sand. S. H. GRAF AND R. H. JOHNSON. *Proc Am Soc Testing Materials* 30, Pt. 1, 574-99 (1930) — Thirteen different methods were studied. For total moisture it was found that the oven drying method was most accurate, the gasoline distn. method 2nd and drying to const. weight with denatured alc. 3rd. In simplicity of operation the distn. method is 1st, oven drying 2nd and the alc. method 3rd. From the standpoint of equipment needed the alc. method is 1st and the oven method and the distn. method are about equal. For surface moisture only, the sp. gr. method with $ZnCl_2$, $MgSO_4$, $CaCl_2$, and $NaCl$ rank in the order named. The A. S. T. M. flask is more accurate than the sp. gr. method with $NaCl$. The colorimetric methods using a weak acid soln. of $CuSO_4$ are more accurate than the elec. method, but not accurate enough for general use. In simplicity the sp. gr. and A. S. T. M. flask methods are about equal. The alc. method is adapted for use where app. for the distn. or oven method is not available. To det. the point at which the sand is surface dry, the Pearson glass jar method is most simple and is satisfactory enough for control work. Everything considered the tests appear to favor the A. S. T. M. flask. The various methods are briefly described. H. C. PARISH

Standardized sieving methods. ALAN PROBERT. *Eng. Mining J.* 131, 311-2 (1931) — Lack of uniformity in sieving data has always handicapped milling operations. A method is described which promises consistent, reliable results and easy and rapid manipulation. W. H. B.

Economics of recovering by-product carbon dioxide. C. L. JONES. *Ind. Eng. Chem.* 23, 519-23 (1931) — The evaluation of by-product CO_2 as a raw material for the manuf. of solid CO_2 is governed by the following factors: purity, pressure, location, seasonal value, safety factor and time value. R. D. BUMBACHER

Analyses of smoke gases. KARL WOLLIV. *Die Gasmaste* 3, 1-5 (Feb., 1931) — In connection with the development of gas masks suitable for firemen a large no. of smokes likely to be met in fighting fires were analyzed for their gas contents. Some smoke samples were prepd. under lab. conditions and others were taken from industrial plants and actual fires. Air mixed with blast furnace gas varied in CO content from about 1% to about 33%, the latter being practically pure blast furnace gas. Smokes from cellar fires varied in CO content from 0.04 to 0.4%, in CO_2 content from 0.03 to 3.4%, and in O_2 content from 17 to 21%. Smokes from roof fires varied in CO content from a trace to 0.2%, in CO_2 content from 0.1 to 2.7%, and in O_2 content from 17.7 to 20.7%. Smokes from miscellaneous fires varied in CO content from 0.01 to 1.4%, in CO_2 from 0.0 to 10.1%, and in O_2 from 9.9 to 20.8%. There was, of course, a wide variation in the compn. of smokes produced under exptl. conditions. A. L. KIBLER

Poison gases in industry and in fires. E. SMOLCZYK. *Die Gasmaste* 3, 9-12 (Feb., 1931). — A discussion from the standpoint of the protection of firemen. A. L. KIBLER

Furfural as an industrial poison and its determination in air. I. M. KORENMAN AND J. B. RESNIK. *Arch. Hyg.* 104, 344-56 (1930) — The presence of small concns. of furfural in the air may be detected by the red coloration produced upon strips of paper treated with a reagent consisting of a mixt. of equal vols. of CaH_2NH_2 and 80% $AcOH$. A quant. est. may be made by passing the air through H_2O contg. the reagent and comparing the resulting color with that produced by a known concn. of furfural. E. R. MAIN

Antifreeze methanol hazard. MAX TRUMPER. *International Clinics* [41], 1, 85-9 (1931) — From a crit. study of the literature, the conclusion is drawn that "a max. permissible concn. of methanol for this new use in the radiators of automobiles should be established, such concn. being detd. by thorough studies made with reference to all types of persons likely to be exposed to its vapors. The general public should not constitute an exptl. lab. for the testing out of a hazard new for the sake of a new industry." A bibliography is appended. JOSEPH S. HEPBURN

Methanol antifreeze and methanol poisoning. W. P. YANT, H. H. SCHIRRENK AND R. R. SAYERS. *Ind. Eng. Chem.* 23, 551-5 (1931) — A status report on an investigation to det. the hazards accompanying the use of methanol as an antifreeze. Expts. limited to breathing of vapors and absorption of liquid through the skin indicate that there is no hazard to health from the reasonable use of methanol for antifreeze purposes. Warming agents to protect against the use of methanol for beverage purposes are recommended. A. L. KIBLER

Ocular hygiene in industry. V. DR LINDA. *Folia Med.* 11, 200-2 (1930). *Bull. Hyg.* 6, 201 (1931). A brief description of interest to ophthalmologists in industry. Effect of lighting, certain irritating gases from chemical factories, dust from metal or stone, intoxication from Hg, P, nitrobenzene and As are described. The preventive aspect is barely touched. GEORGE R. GREENBAUM.

Commercial and war gases and the chemistry of protection against them. J. LINDA. *Chem. Rundschau Mitteleurop. Balkan* 7, No. 15, 28-33 (1930).—A detailed account of different noxious gases, of methods of preventing their toxic action and of the composition of gas mask filters. S. S. DE FINKEL.

Behavior of the catalyst in the canister of the carbon monoxide gas mask. G. STAMPI AND J. R. HANCIET. *Z. ges. Naturwiss. Springer* 26, 246 (1931).—Gas mask canisters for all gases except CO depend upon the principle of physical or chemical sorption and their useful life is determined by the amount of gas absorbed. In the CO mask, however, the amount of CO absorbed plays an important part in its useful life, which depends in large measure upon the other foreign gases in the atmosphere. It is especially sensitive to moisture and for this reason must contain a very efficient drying agent. S. and H. describe the behavior of the CO canister when other poison gases, both alone and in addition to CO are present in the air. With an air flow of 30 l. per min. and a toxic concentration of 1 vol.-%, the amounts of toxic gases in g. taken up by the Dräger CO canister and the breakpoints in min. are as follows: phosgene 3.4 g., 27 min.; Cl₂ 7.2 g., 81 min.; HCN, 3.3 g., 98 min.; SO₂ 18.5 g., 232 min.; H₂S, 19.9 g., 467 min.; NH₃ 2.3 g., 108 min.; CCl₄ 8 g., 41 min.; C₂H₆ 10 g., 103 min. The effect of other gases in addition to CO was tested by passing a continuous air flow of 30 l. min. through the canister. During the first hr. of the tests CO and another toxic gas were added to the air stream. These were then omitted for 1 hr. and added again for another hr., etc. The concentration of CO was 0.5 vol.-% while the concentration of the other gas used varied. The relative humidity was 30% in all cases but one, when 90% was used. The breakpoints for the second toxic gas were determined and the points of maximum allowable concentration in the exit gas for CO. The results are shown by curves from which the following points are taken:

Second gas added	Breakpoint (v. second gas, total hrs.)	Max. allowable concn. CO, total hrs.
None		CO
0.03 vol.-% H ₂ S	15.5	24
0.026 vol.-% HCN	10.75	19
0.031 vol.-% SO ₂	21	15.5
0.015 vol.-% Cl ₂	13.5	14
* None, but with 90% relative humidity		11

The catalyst (Hopalite) is poisoned to an equal extent by the following parts of the gases named: Cl₂ 10, SO₂ 23, HCN 24, H₂S 30, H₂O vapor 150 (drying agent present). Condens. shown on curves are 10 times those shown in the text and are probably in error.

Misleading propaganda on refrigerants. J. R. CHURCHILL. *Refrigerating Eng.* 21, 204-71 (1931).—C. replies to a report of the committee on poisonous gases of the American Medical Association in which the use of methyl chloride as a refrigerant is discussed. A. H. JOHNSON.

New applications of low-temperature freezing. FRANK ZIMMERO. *Refrigerating Eng.* 21, 251-64 (1931).—The application of the quick freezing process to peaches and ice cream is discussed. Temps. between -15° and -35° F. have been found suitable, but in general the more rapidly the peaches can be frozen or the ice cream hardened the more satisfactory will be the results. The refrigerant used for obtaining the low temp. is usually CO₂. However, by combining an NH₃ and a CO₂ cycle, a combination is obtained whereby the very high pressures commonly used in CO₂ plants are eliminated. A. H. JOHNSON.

Problems of insulation. F. MOLL. *Korrosion Metallschutz* 7, No. 2, 39-41 (1931).—Insulation of metals against the access of moisture by means of impregnated cellulose fibers has not proved entirely successful in that the fiber in spite of impregnation picks up water. Cables are found puffed because of corrosion due to the presence of water. The absorption of water by the fibers is attributed to the presence of a unimol. layer of water on the crystal faces at the intermol. boundaries. The presence of this layer is indicated in that the quantity of water absorbed by a unit quantity of cellulose when exposed to a humid atm. corresponds exactly to the quantity of water estimated from the dimensions of the mols. and the assumption of a unimol. water layer.

R. E. ROTHFELD.

Cable and other insulators. RUDOLF DITMAR. *Kolloid-Z* 54, 355-7 (1931) —
 Patent review ARTIUR FLEISCHER

Electrical materials commonly used—paper (HIRSCH) 23. De-termination of NH_3 in refrigerator brines (GIBBS) 7. Radiation pyrometry (FRIDGE) 2. Mechanical refrigeration in the sugar industry (DOMITSKY) 28.

BADGER, W. L. Heat Transfer and Crystallization. Harvey, Ill. Swenson Evaporator Co. 22 pp. Reviewed in *Kodak Research Labr. Monthly Abstract Bull* 17, 172 1931

Bibliographie des livres français sur l'industrie et la technologie, 1919-1930. Paris. Hotel du Cercle de la Librairie. 234 pp.

BLICHER HANS AND LANGE OTTO. *Ankunftsbuch für die chemische Industrie*. 14th ed. revised by O. Lange. Berlin: de Gruyter. 592 pp. Lnen. M 50

Grafes Handbuch der organischen Warenkunde. Edited by VICTOR GRAFE. Band III. Halbband I. Die Gärungs-gewerbe. Ruchstoffe. Organische Säuren. Natürliche Farb- und Gerbstoffe. 708 pp. M 49.50, half lnen. M 53. Band IV. Halbband I. Narkotische Genussmittel. Drogen. Gewürze. Harze. 688 pp. M 44.50, half lnen. M 48. Stuttgart: C. E. Pöschel Verlag. Cf C A 25, 1925

Mückenbergers Handbuch der chemischen Industrie der ausserdeutschen Länder, 1931 32. 7th ed. revised. Arranged for use in German, French and English. Berlin: Urban & Schwarzenberg

RASSOW E AND LOESCHKE A. Jahresbericht über die Leistungen der chemischen Technologie für das Jahr 1930. Abt. I. Unorganischer Teil. 70th year. Leipzig: J. A. Barth. 781 pp. M 46 bound, M 49

Estimating the hygroscopicity of materials in factories and warehouses. THOMA & Co. Swiss 142,491, Nov 23, 1929. A sample of the material being manufd. or stored, e. g., cotton, is tested in a balance, and from these data the hygroscopicity of the treated material is detd.

Heating processes. ALLOEMEINE ELEKTRIZITÄTS-GESELLSCHAFT. Ger 521,139, July 23, 1925. Heat is conveyed to smelting furnaces, welding app., etc., by passing a di- or poly-atomic gas, preferably H_2 through an elec. arc or other heating means capable of dissociating a substantial part of the gas, and leading the gas to the app. to be heated, where the atoms recombine to molh., with liberation of heat. Addnl. details are given and app. is described

Catalytic reactions. ALFRED REIS. Fr 699,131, May 29, 1930. In org. catalytic reactions, the contact masses are in intimate contact with heat-conducting metallic surfaces which are placed in the contact chamber parallel to the direction of flow of the gas.

Electrothermal reactions. BUFFALO ELECTRIC FURNACE CORP. Fr. 697,103, June 5, 1929. Reactions of solid powders at high temps. are carried out by passing the materials through a hollow electrode from the end opposite to the arc in a furnace. The electrode is maintained at a high temp. so that the reaction is complete before the materials reach the end of the electrode

Separating gas mixtures. AMERICAN SMELTING & REFINING CO. Fr 697,055, June 2, 1929. A mixt. of gases, such as gases from metallurgical app., is dried and sep'd. by passing the mixt. through an adsorbent material sat'd. with one of the constituent elements and allowing the water contained in the mixt. to replace the constitutive element in question in the adsorbent material. An app. is described. Cf C A 25, 2105

Separating gases by liquefaction. ARTHUR SELIGMANN. Ger 520,287, Feb. 20, 1930. A method and app. are described for removing residues of liquefied gas from the plant when the latter is shut down.

Removing phosphine from gases. SOC. D'ÉTUDES SCIENTIFIQUES & D'ENTREPRISES INDUSTRIELLES. Fr 696,512, May 21, 1929. PH_3 is removed from gases contg. it in small quantities, e. g., H_2 obtained by the reaction at high temp. and pressure of P with water, by treating the gas at a high temp. with O or an oxidizing gas. If air is used in suitable amt. a gas contg. H and N in the proportions for making NH_3 may be obtained.

Recovering inflammable solvents. JEAN H. BRÉGEAT. Ger 520,076, Aug. 28, 1928. Vaporized inflammable solvents are condensed by bringing the vapors into direct contact with blocks of compressed CO_2 snow. The gaseous CO_2 thus liberated acts as a protective gas against fire

Colloidal solutions. HANNS STAMMEICH and WILHELM NOHER. Ger. 521,203, Dec. 25, 1927. In prepg. colloidal solns. with the use of protective colloids, the ma-

terial to be colloidized and the protective colloid are produced simultaneously from one soln or liquid. Thus, a colloidal soln of AgCl and SiO_2 may be prepd by adding HCl to a soln of AgNO_3 and water glass, or an organosol of Ag and C may be prepd by passing an air through a suspension of a Ag salt in paraffin, or through paraffin between Ag electrodes.

Stabilization of solutions, emulsions and suspensions by lignin derivatives. CHAS F CROSS and ALF ENGELSTAD. Ger 516,566, Sept 18, 1926. See Brit 263,520 (C A 22, 162).

Automatically controlling density or character of solutions or mixtures by differences in their light transmission. WM A DARRAU. U S reissue 18,003, March 21. A beam of light passed successively through the material to be controlled and a standard may serve to control dye solns. operations such as bleaching and filtering of sugar solutions or regulation of the d of wood pulp aq mixts. Numerous details and elec features are described. This pat is a reissue of the original pat. No 1,716,523, issued Feb 11 1930.

Purification of liquids by distillation. GEORGES L GARDEL. Fr 695,266, May 7, 1930. The mass of liquid to be purified is first heated to a temp near the lowest b p of the components and an inert gas or dry steam is introduced under high pressure. The gas is adsorbed and by expansion produces a lowering of the temp while at the same time carrying over the mols. A suitable app is described.

Liquefying chlorine. KRUMHOLTZ & CO, G M B H. Ger 514,070, April 6, 1929. The condenser comprises a no. of practically horizontal tubes arranged one above another and connected at the ends by vertical mains. The gas is supplied at about the middle of one main, and the liquefied Cl is collected at the bottom of the other main, the unliquefied impurities escaping up the outflow main.

Removing liquid from gelatinous colloid emulsions. WILFRED S ROTHBERG, STANLEY BLYTHEN and H K GILLSPER (Kurt Ripper, inventor). Ger 517,731, July 28, 1923. For producing hardened objects, emulsion colloid is set in molds permeable to heat, and formed of porous walls which allow passage of the liquid, but not of the colloid. The molds may be heated. Cf C A 19, 1464.

Separating mixtures of liquids. N-V DE NATAAFSCHIE PETROLIUM MAATSCHAPPIJ. Fr 696,324, May 30, 1930. A liquid mixt is sep'd into its constituents or groups of constituents by submitting the mixt to the action of 2 liquids circulating in counter-current and which are very insol in one another, these auxiliary liquids being so chosen that the constituents or group of constituents are divided between them in different proportions. The auxiliary liquids have preferably different sp wts and one of the liquids may be miscible with the mixt to be sep'd.

Separating mixtures of materials of various degrees of granularity by sedimentation. CHEMISCHES LABORATORIUM FÜR TONINDUSTRIE UND TONINDUSTRIE ZERTUNG H SEYER & E CRAMER G M B H. Ger 517,960, July 26, 1929. Details of the heating and agitating of the sedimentation liquid, and of the app, are given.

Lixiviating materials. GEORG WEISHERT. Ger 517,920, June 23, 1927. App for lixiviating materials, especially ores contg nitrates or chlorides under heat and pressure, is described. Countercurrents of lixiviating liquid are forced through the heated material.

Substances in uniform particles. I G FARBENFABR A-G. Fr 696,096, May 26, 1930. Fusible org or inorg substances are transformed into particles of uniform size by dividing the fused substances at a temp near the solidification pt by means of rotating disks or centrifuges, regulating the speed of rotation so that spherical particles are obtained and solidifying by cooling with a gas or liquid.

Pulverizing vegetable matter. SOC ANON FOUR L'IND CHIM A BALE. Swiss 143,521, Apr 11, 1928. The method is similar to those of 128,835 and 139,321 (C A 25, 776), i e, the temp is lowered to cause the material to congeal, after which it is pulverized. This method facilitates good extr.

Drying crystalline or other materials. GUTHROFFNUNGSHÜTTE OBERHAUSEN A-G. Brit 340,057, Oct 8, 1923. Materials which may be in cryst dust like, granular or lump form or in blocks or other shaped masses are dried by making them resistances in the circuit of a high frequency elec current, and the drying may be assisted by a stream of gas.

Refrigerators. SIMMONS-SCHUCKERTWERKE A-G. Fr 696,118, May 26, 1930. In an absorption refrigerator, a working agent is used which boils at atm pressure in the neighborhood of 100° and the m p of which is below the evap temp entering into account for the production of cold. A soln of ethyl isomyl ether in paraffin oil or a soln of toluene in quinoline may be used.

Refrigerating agents. COMPAGNIE FRANÇAISE POUR L'EXPLOITATION DES PROCÉDÉS THOMSON HOUTON Fr 695,546, May 29, 1930 HCCOMe is used as a cooling agent, particularly in refrigerators

Refrigerating agents. I RIGIDAIRE CORP Fr 696,543, May 28, 1930 An absorbent material for refrigerating agents is composed of a mixt of a solid (such as SrCl_2) capable of forming a chem combination with the refrigerating agent and LiNO_3 , the latter being 4-10% preferably 5% of the total mixt and being in the liquid state under working conditions

Foam prevention during refrigerating. A BORSIO G M B H (Gustav Jansen, inventor) Ger 520 120 July 11, 1929 Foaming of the brine is prevented by addition of oil

Heat insulation. PAUL E. KLOPSTIG (to Central Scientific Co) U S 1,797,265, March 24 Heat resisting walls suitable for still jackets, etc., comprise a molded thick porous layer of heat insulating material such as diatomaceous earth 75, hydrated lime 12% and asbestos fiber 12% and a molded facing layer of dense heat resisting material such as diatomaceous earth 75 and portland cement 25%, the 2 layers being molded one on the other and having about the same coeff of expansion

Protection and heat insulation of iron tubes. RUDOLF HIRSCHER Ger 521,077, Nov 13, 1927 The tubes are thinly coated with nitrocellulose or like cellulosic material, dipped in a hot bath of asphalt bitumen or tar, and then rolled under pressure in powdered asbestos, talc or kieselguhr

Electric insulating compositions. I G FARNEVIND A G. (Max Hagedorn, inventor) Ger 519 457, Sept 27, 1927 The compns consist of (or) contain cellulose esters of higher fatty acids, or mixed esters or ether esters contg higher fatty acid radicals

Thermal insulation for electrolytic cells. JAMES N SMITH (to Westinghouse Elec. & Mfg Co) U S 1,797,374 March 21 Structural features

Impregnating fibrous insulation, etc. STANDARD TELEPHONES & CABLES, LTD, T R SCOTT and T N RILEY Brit 119 911, Sept 18, 1929 In operations such as the manuf of high tension cables, elec condensers, etc., the insulation is dried and exhausted, and, before impregnation, residual air is replaced by the vapor of a substance which vaporizes under the temp and low pressure conditions prevailing in the impregnating vessel and which is condensed to a liquid or solid under the conditions of temp and pressure in service The final slushing of a cable may serve as the impregnating vessel and preferably the substance used is such as to condense during impregnation Among the materials which may be impregnated are paper, textile materials, fiber boards, asbestos and wood, and as impregnating substances there may be used, hydrocarbon oil with or without resins, tars, bitumens, resins and synthetic resins in soln, cellulose acetate solns, S, paraffin, ceresin, beeswax, hydrogenated waxes such as chlorinated naphthalene and petroleum jellies Methyl-naphthalene and biphenyl also may be used They are stated to reduce the dielec losses in the finished cable Numerous details of procedure are described, and substances of high sp inductive capacity such as glycol or ethylene glycol may be used in condenser manuf

14—WATER, SEWAGE AND SANITATION

EDWARD BARTOW

Developments in western municipal water-supply practice. HARRY N. JENES, *Eng News Record* 106, 414-40(1931)—Consideration of existing supplies from distant mountain watersheds leads to the conclusion that it may be less expensive and more reliable to filter the water than to provide the increasingly elaborate watershed protection now demanded Practically all mountain sources, through either natural or artificial causes, are becoming less desirable in regard to quality and in recent years the water from practically every source, even under apparently ideal conditions of mountain catchment area, has been subjected to some form of treatment In California, the most striking development is the reclamation of water from municipal sewage and storm water drainage Some of the salient characteristics of western municipal water supplies and comparative cost statistics are tabulated R. E. THOMSON

Ground water in eastern and central Montana. EUGENE S. PERRY, *Mont. Bur. Mines and Geol Mem* No 2(1931)—A report on the occurrence of artesian water in Montana, the effect of geological structure upon this occurrence, and the chem compn of various types of ground water CURTIS L. WILSON

Operating conditions at Elbe Waterworks and pretreatment of Altona ground water supply. LICHTENH. *Gas u. Wasserfach* 74, 237-44(1931) —A description of the pumping plant and rapid filter is given with special reference to difficulties encountered in the very cold January of 1929. Acidified (H_2SO_4) clay suspensions are added to water before rapid filtration to improve filtration. The deep wells are described and illustrated. R. W. RYAN

Oberhaus water works with quick filter plant and infiltration plant for augmenting ground water. HEINRICH KRING. *Gas u. Wasserfach* 74, 193-9(1931) —Costs of operation are given for filter and infiltration plants and details of operations. Illustrated. R. W. RYAN

Chemical and physicochemical investigation of the sulfur-bearing mineral waters of St. George in Angarano. (Bassano del Grappa). G. BRAGAGNOLI. *Ann. chim. applicata* 21, 12-26(1931) —An extensive examn of these waters was made indicating the presence of 23 cc of H_2S per l of water at 0° with 0.64 g solids (mostly Ca and $\text{Mg}(\text{HCO}_3)_2$). The therapeutic value of the waters was not investigated. A. W. C.

Gravity water-supply in Franklin, Tenn. PARK MARSHALL. *Am. City* 42, No. 3, 7(1930) —Well water brought 12 miles and stored in a reservoir and in an elevated tank is chlorinated. C. H. BADGER

Automatic chlorination of a gravity water-supply. PERCY R. SANDERS. *Am. City* 42, No. 3, 94-5(1930), cf. C. A. 24, 1445, 5397. C. H. BADGER

The chlorine-binding capacity and the potassium permanganate requirement of water and sewage. E. BARTH. *Arch. Hyg.* 104, 318-29(1930) —Addns of uric acid and sugar increase the Cl no. of boiled tap water, but do not affect the Cl requirement. Both are increased to a greater degree by the addn of ferrous salts or nitrites than by ferric salts or nitrates. Cl in a concn of 1 mg per l has a marked bactericidal effect on tap water, the bacteria requiring less Cl than an equal wt of org substance. A concn of 20 mg per l is required to produce a marked decrease in the bacterial count of sewage. E. R. MAIN

The analysis of water. V. MADARA. *Chem. Listy* 25, 13-14(1931) —A critical review of analytical methods used in H_2O analysis. FRANK MARESH

Improved water sample bottle. ROSS L. LAYBOURN. *Am. J. Pub. Health* 21, 419-20(1931) —Glass stoppered, metal capped bottles of a type which can be purchased from a physician's supply house are used. J. A. KENNEDY

Mixing basins—time of mixing. JOHN R. BAYLIS. *Water Works and Sewerage* 77, 279(1930) —Mixing periods in filtration plants vary from 3 to 90 min but generally the mixing time in newer plants is 30 to 60 min. Factors that influence time of floc formation are quantity of coagulant, time of mixing, temp., violence of agitation and character of water. C. C. RUEHMORT

Odor and taste control at Charleston, W. Va. LAWSON HAYNES. *Water Works and Sewerage* 77, 266(1930) —See C. A. 25, 1308. C. C. RUEHMORT

The Kempton Park Primary Filters of the Metropolitan Water Board. S. WALKER. *Water & Water Eng.* 32, 562-7(1930) —A general description of the construction of the plant. EDWARD S. HOPKINS

Prechlorination of filters. JOHN F. KINGSLEY. *Water Works and Sewerage* 77, 194-5(1930) —Prechlorination of the raw water at Newburgh, N. Y., has accomplished several beneficial results without any significant increase in cost of operation. Prechlorination eliminates short filter runs, reduces the size and the number of the mud balls, eliminates surface cracks and wall cracks in the sand bed and removes the color more readily without increasing the coagulant dose. High velocity washes are not necessary when water is prechlorinated. C. C. RUEHMORT

Washing rapid sand filters. ANON. *Public Works* 61, 49-51(1930) —Studies were made over a period of 2 years at the Detroit Experimental Filter Plant. Shrinkage of filter sand bed and other attendant filter troubles were overcome by removing the coating from the sand grains and the mud accumulation from the beds. By washing the filters at a sufficient rate to obtain 50% or more expansion of the sand coating and shrinkage were prevented permanently. A method for cleaning dirty filters is presented. It requires the use of NaOH and hosing with a $\frac{1}{2}$ -inch nozzle. A sand-expansion index for the filter wash is recommended to supplant the generally used index of rate of wash water rise. C. C. RUEHMORT

Cold-weather troubles in filtration. C. T. HOUGH. *Water Works and Sewerage* 74, 260(1930) —See C. A. 25, 1308. C. C. RUEHMORT

Factors influencing the toughness of coagulated matter. JOHN R. BAYLIS. *Water Works and Sewerage* 77, 184-8(1930) —Microorganisms have the greatest influence in shortening filter runs. Gelatinous ppts of org origin influence the strength of coagu-

lated matter Increase of temp increases the strength but it is not desirable to increase the temp of the water Changes in chem treatment may produce a stronger flocculation The flocculated turbidity must be very low in the water going to the filter beds when the temp is low to prevent flocculated matter from passing the beds, unless there is something in the water that adds toughness to the flocculated matter

C. C. RUCHHOFF

Ferric iron coagulation in a plant of unique design. ALEXANDER POTTER AND WAR I. KELLIN *Water Works Eng* 83, 949-50, 970-3(1930) —The treatment given at the Westchester Joint Water Works No 1 is described in detail (cf *C. A.* 24, 5902) The best coagulating results are obtained by mixing not in excess of 10% of the theoretical required amt of dry Cl_2 gas with the iron sulfate soln and applying to the prechlorinated raw water The superiority of ferric iron coagulant over $\text{Al}_2(\text{SO}_4)_3$ coagulant for this particular water is that ferric iron gives a heavy floc formation with quicker settling, even with low temp and over a wider range of pH , resulting in better plant operation Also in *J. Im. Water Works Assoc* 23, 719-27(1931) C. H. BADGER

Sterilization of water mains. CLAY P. STEWART *Water Works and Sewerage* 77, 191(1930) —Whenever water mains are repaired or new pipes installed liquid chlorine is used to sterilize the section of the main that may possibly have become contaminated Portable chlorinators have proved satisfactory for this work C. C. R.

Chemical and biological study of deposits on the interior of water pipes. R. KOLKOWITZ AND E. BEGER *Gas u. Wasserfach* 74, 267-8(1931) —Wood, glass and metal plates were exposed to tap water by means of a special device The plates were then examined microscopically, permitting a convenient biol. analysis of the water similar to that obtained by sieving The nature of pipe deposits is discussed R. W. R.

Conditioning refinery water. SMNEY BORN *Natl. Petroleum News* 23, No 6, 52-3, 59-7, No 7 55-57, 58-60, 61-99 No 8, 51-1 54-57, No 9, 63-8, 70 No 10, 49-50, 52, No 11 70-1(1931) —The following topics are discussed: impurities found in natural waters; methods and app. for testing for hardness, alk., causticity, acidity, salt and total dissolved solids; the benzidine method for detg. sulfates in water; limits of hardness, alk., permissible in general purpose waters; definition of pH value, chem. reactions involved in water treatment, boiler feed water, loss of heating efficiency due to scale, removal of scale-forming materials, foaming and priming, embrittlement prevention, zeolite treating; chemistry of process, operating precautions, construction and operation of soda lime water softeners, intermittent and continuous types, hot and cold processes; reduction of scale-forming materials; prevention of corrosion, comparison of zeolite and chem. type softeners, choosing the softener, boiler compds., water for cooling and condensing; drinking water at industrial plants and oil camps, filtration; Cl_2 , violet ray and ozone treatment, water for treating purposes R. W. KELLY

How contamination from pump packing was checked. ARTHUR F. MELLEW. *Water Works Eng* 83, 1040-3(1930), cf *C. A.* 24, 4879 C. H. BADGER

Maintaining chemical balance to resist corrosion and its applications to a recent development in lime-water softening. CHARLES P. HOOVER *W. Va. Univ. Eng. Expt. Sta., Tech. Bull.* No 3, 17-24(1930) —See *C. A.* 24, 3583 G. G.

Maintaining chemical balance to resist corrosion. CHARLES P. HOOVER *Water Works and Sewerage* 77, 287(1930) —Remedies for preventing corrosion are: (1) Water should be preheated to drive off oxygen (2) It should contain 30-50 p. p. m. of alkalis in excess of that required to soften it. The higher the pH the smaller the amount of iron dissolved (3) Water can be passed through a bed of scrap iron or filings (4) Pipes may be coated with zinc, cement or tar (5) Twenty p. p. m. of sodium silicate is sufficient to stop corrosive action almost entirely (6) Water may be so treated that it will deposit an impervious film on the inside of a pipe This film prevents direct contact with iron The film can be built up by satg. the water with CaCO_3

C. C. RUCHHOFF

Corrosion and conservation of underground structures. P. J. RICHARDS. *J. Am. Water Works Assoc.* 23, 529-33(1931). —See *C. A.* 25, 1609 H. G.

Correcting corrosive water at institution. G. M. RIDENOUR. *Public Works* 61, 122-3(1930) —A New Jersey state institution reports a persistent and serious case of corrosive water The water supply is secured from a 260 foot well driven in yellow sand, blue sand and clay Analysis at the well gave: pH 5.5, CO_2 acidity 14.0 p. p. m., total alk. as bicarbonate 10.0 p. p. m., dissolved O 2.00 p. p. m. and hardness 22 p. p. m. The tap water analysis gave, in addn. to the above, 4.9 p. p. m. of ferric iron Installation of a deaerator reduced the acidity to 2.0 p. p. m. and increased the pH to 6.2, but failed to check the corrosive action The VanHeuser marble test was used to det.

a point of equilibrium between the carbonate, bicarbonate and CO_2 such that a slight coating of carbonate would be deposited on pipe surfaces and thus serve as a protection from O or acids. Lime was applied at the pump suction in sufficient quantities to maintain a pH of 7.7 to 8.0. This treatment resulted in a CO_2 acidity of 0.2 p.p.m., a bicarbonate alkali of 43 p.p.m. and an occasional carbonate alkali up to 8 p.p.m. Conditions were markedly improved after 2 months of the lime treatment. Approx. 30 lb. of lime is used per 100,000 gallons of water. C. C. RUCINOFF

Some recent developments in boiler-water treatment. A. W. CHAPMAN. *Fuel in Science & Practice* 10, 64(1931).—The disadvantages of boiler scale, its formation and prevention are reviewed. D. A. REYNOLDS

Ways of representing analytical results of the examination of boiler feed water. J. LEICK. *Z. anorg. Chem.* 44, 100-2(1931).—A general article comparing the methods of expressing the quantitative results of water analysis. RUSSELL C. FAN

Industrial sewage disposal. S. JACKSON and W. GARNER. *Ind. Chemist* 7, 76-9, 101-4, 166-8(1931).—A review. E. H.

The present trend in sewerage and sewage-disposal practice. W. W. HOANER. *J. Eng. Club St. Louis* 10, No. 9, 12 pp. (1930). E. I. S.

Future possibilities in sewage disposal. H. C. SHERTON. *Surveyor* 79, 381(1931).—An ideal system would consist of a preliminary settling tank, activated sludge tanks in one or two stages, digestion tanks for mixed sludge, arrangements for heating and mixing sludge keeping the temp. at 75-80° F., covered sludge-drying beds, and a power plant driven by the gas evolved. A. L. ELDER

Sewage-disposal works of Plainfield, N. J. JOHN R. DOWNES. *Water Works and Sewerage* 77, 189-91(1930).—This plant is designed to treat 4 m.g.d. There are 6 Imhoff tanks, 2 sprinkling filters, 1 final settling tank and 5 sludge-drying beds. Provisions are also made for partial aeration of the sedimentation tank effluent in which Cl₂ is used for odor control. C. C. RUCINOFF

Sewage treatment at Lake City, Fla. ANON. *Public Works* 61, 110(1930).—Lake City, Florida, with a population 3000 and an average daily flow of 100 gallons per capita per day has recently constructed a sewage treatment plant having the following units: preliminary bar screen, a primary clarifier, pump house with wet and dry well, dosing tank, sprinkling filter, sludge digestion tank, secondary clarifier and sludge-drying bed. Provision is also made for pre- or post-chlorination of sewage. C. C. R.

Los Angeles successfully reclaims sewage for replenishment of underground water supplies. C. WILSON. *West Construction News* 5, 473-4(1930).—Discussion of sewage-disposal conditions in southern California with special reference to the program of water reclamation studies initiated by the Los Angeles Dept. of Water and Power. E. I. S.

Sewage screening—experience with disk-type screens at Rochester, N. Y. WILLIAM A. RYAN. *Am. City* 42, No. 3, 101-2(1930).—Sewage treatment was begun in 1915. The city is now served by 4 plants. The largest, the Irondequoit plant, built in 1916, was designed to treat a normal dry weather sewage flow of 24 m.g.d. The flow, since then, has increased to almost 50 m.g.d. without causing the operating difficulties usually experienced. This is attributed to the fine screening installation. The primary purpose of these screens is not to remove the solids which will settle out in the Imhoff tanks but to prevent the floating solids from entering these units or being discharged under storm flow conditions into Lake Ontario. The addition of 2 new disk-type screens in 1923, making a total of 6, now permits handling a rate of 180 m.g.d. without by-passing crude sewage into the lake. A feature of the new screens is the variable speed transmission unit whereby the lineal travel of the screen may be varied instantly from 7 to 27 ft. per min. to meet existing flow conditions. The equipment of the plant is listed. The effluent is discharged through a 66-in. steel conduit and submerged outlet 7000 ft. offshore into the lake. C. H. BADGER

Plans for sewage reclamation in the Los Angeles metropolitan area. R. F. GOUDEY. *Eng. News-Record* 106, 443-6(1931).—See C. A. 25, 1928. R. L. THOMPSON

Sewage treatment in Germany. SENTENAC and FONTAINE. *Ann. ponts et chaussées Pt. Tech.* 100, No. 3, 266-97(1930).—Notes on an inspection and study trip with special reference to the methods and equipment observed in Berlin and in the Ruhr District. E. I. S.

Sewage purification by dilution and the limit of putrescibility. H. BACH. *Techn. Gemeindeblatt* 34, 29(1931).—The amount of water needed for dilution depends on the strength of raw sewage. This factor varies greatly in different countries. The ratio of material in solution to that in suspension is an important factor. A. L. ELDER

Some experiments on the treatment of a sewage containing wool-scouring refuse.

C. C. BEEBHAM, *Surveyor* 79, 335(1931)—The sewage of Bradford contains large quantities of wool scouring liquor. Raw effluents have analysis such as: total solids 42,000 p. p. m., grease 7900 p. p. m., alkyl to methyl orange 31,200. Good purification was obtained by neutralization of the alkyl and mixing gently with a little fresh sludge, settling and filtering at 60° F. A. L. FIDLER

The question of gas formation in city sewers. KRIEGER AND MCILPHER, *Techn. Gemein. Abfall* 33, 303 9(1931)—Cases of explosions in sewers are cited. The limits of CH_4 , CO , etc. needed to produce explosive mixtures are given as well as the toxic contents. Discussions of methods of testing for explosive and poisonous gases in sewers are shown. A. L. FIDLER

Purification of brewery sewage by the activated-sludge process. M. A. ILL, *Techn. Gemein. Abfall* 33, 291 5, 309 12(1930). The activated sludge process was a satisfactory method for treating brewery waste mixed with other sewage. The pH needed to be controlled. A. L. FIDLER

Influence of seeding material on sludge digestion. C. I. KIEFER AND HERMAN KRATZ, JR., *Eng. News-Record* 106, 474 8(1931). The following conclusions are drawn from lab. expts. conducted at the Baltimore sewage works: (1) Each g. of volatile matter destroyed or lost as CO_2 by present dry testing methods, will be converted into about 750 to 800 cc. of CH_4 , and 1,200 to 1,900 cc. of total gas. Each g. of raw volatile matter digested will produce about 400 cc. of CH_4 , and 800 cc. of total gas. (2) Sludge that is several years old serves just as well for seeding purposes as sludge that has just reached a digested condition. (3) Partly digested sludge, even though it contains as much as 65-70% volatile matter, becomes digested within about 60 days when 2% daily oblique of raw material are made, and within 70-80 days when 6% of raw sludge is added daily. When raw sludge, in place of digested or partly digested sludge, is used, 150 days are required to obtain ripe material in the case of 2% additions. Where 6% additions were made, a digested sludge was never obtained. (4) The sludges receiving 2% additions were slightly better digested than those to which 6% was added. The difference, however, was not great. It is believed that 6% additions are about the max. which can be made and still produce a sludge with good drainability, free from odors. (5) Data obtained indicate that certain ether-sol. materials digest rapidly within 10-15 days. There is present, however, ether-sol. matter that digests at a slow rate. R. L. THOMSON

Stream pollution by textile wastes. RICHARD WESTON, *Public Works* 61, 110 (1930).—See C. A. 24, 450. C. C. RICHMOND

The effect of certain poisons upon mosquito larvae. MARTIN FROMMELT, JR., AND RAYMOND C. SHANNON, *Am. J. Hyg.* 13, 611-22(1931).—Full grown larvae of *Aedes aegypti*, *An. leucorhynchus* and *Culex quinquefasciatus* were not injured by relatively high concns. of KCN, but weak solns. of I (1 part per million) were fatal. Newly hatched larvae were much more susceptible. The toxicity of Br was practically the same, while that of Cl was less. Because of the thickly sclerotized cuticle the pupae were more resistant to the halogens. In concns. toxic to larvae I or Br cannot be tasted but their employment in drinking water is questioned. In cisterns, tanks, still water, etc. the larvae are readily destroyed. G. H. W. LUCAS

Olean City (N. Y.) epidemic of typhoid fever in 1928. ARCHIBALD S. DEAN, *Am. J. Pub. Health* 21, 390-402(1931).—This is a report of an epidemic of 218 cases due to a break in the submerged suction pipeline from a well in the Allegheny River and insufficient chlorination of the water. There were 25 deaths (C. N. Y. S. Dept. of Health Repts. 49, 104-12 (1928)). In Jan. and Feb. there were over 1000 cases of gastroenteritis and during Aug. and Sept. over 5000 cases. Olean draws its water from 2 sources: Olean Creek and South Olean wells. The former was free from *B. coli*. The latter supply came from 12 or more 6" and 8" tubular wells sunk about 72' deep in the low ground on the north side of the river and from a large open collection and suction well on the south bank of the river. Two of the tubular wells had become submerged within recent years by changes in the course of the river. Water from these wells passed by screw joint iron suction pipes across the river bed to the collection and suction well. After chlorination, instituted by a device which operated by remote control when the elec. pump was started, the water was forced into the distribution system. Raw sewage from nearly half the city entered the river and its tributary above the wells. No daily or regular tests for residual Cl_2 were made. Nine of 16 samples examined in July, Aug. and Sept. gave presumptive tests for *B. coli*. Considerable epidemiological data are given. Details on the cost of this epidemic are given. J. A. KENNEDY

Trisodium phosphate [as water softener and boiler compound] (SNELL) 18.

Purifying water. JOSEF MUCHEA Fr 695,550, May 14, 1930 In the purification of water by electromagnetic radiations, the modification in optical properties—color, absorption power, reflection, refraction, polarizing effects, etc.—are used as an indication of the actual state of the water to be treated and the progress of the treatment is controlled automatically on the basis of the modification produced.

Sterilizing water. JOSEF MUCHEA Fr 696,342, May 31, 1930 The dechlorination of chlorinated water by means of activated charcoal, etc., is combined with a sterilization of the charcoal by submitting the charcoal during dechlorination to the transposition or stirring used for sand filters.

Apparatus for rapid filtration of water. HENRY A. DESREMAUX. Fr 696,982, May 19, 1930

Distillation plant for crude water. AKT GES BROWN, BOVERI & CIE SWISS 143,480, Nov. 8, 1923

Apparatus for deaerating water, etc. WM S ELLIOTT U S 1,797,405, March 24

Acid removal from water. OSCAR RITSCHL Fr 695,754, May 16, 1930 Waters are decalcified by fixation of their free or semi free CO_2 by passing them through a filtering mass composed of broken or calcined natural MgCO_3 or dolomite.

Softening water. I KRUGER A/S Fr 695,065, May 22, 1930 Water to be completely or partially softened is treated with pptg agents and then brought in contact with granular or subdivided material, which is preferably of a similar chem compn to the substance to be pptd, e. g., marble or limestone, whereby pptn is accelerated.

Softening water with base-exchange material. UNITED WATER SOFTENERS, LTD., R T PEMBERTON and H S LAWRENCE Brit 340,091, Nov 7, 1929 To regenerate base-exchange material (in an app various details of which are described), a rigid rod or cartridge of salt is inserted into a tube or pipe and water is passed through the latter to dissolve the salt and carry it to the base-exchange material. Cf C A. 24, 5907

Apparatus for softening water with base-exchange materials. GOV S APPELOAN U S 1,797,708-9, March 24

Scale preventing. OSCAR RITSCHL Fr 695,753, May 16, 1930. The formation of scale in boilers, etc., is prevented by adding to the feed water a neutral salt of NH_4 , such as $(\text{NH}_4)_2\text{SO}_4$ or $(\text{NH}_4)_2\text{CO}_3$, which only decomposes into NH_4OH and free acids after the heating of the water. This acid afterward sets from a part of the bicarbonates the large amt of CO_2 necessary to maintain the Ca bicarbonates contained in the water.

Sewage treatment. WILHEM RUDOLFS (to National Aluminate Corp) U S 1,797,879, March 24 Sol Fe salts such as FeCl_3 or $\text{Fe}(\text{SO}_4)_3$ or FeSO_4 are added in the proportion of 3.5 to 5.0 p p m, in order to facilitate sedimentation. Use of Mn, Ni and Co salts is also mentioned.

Clarifying apparatus for waste water. DEUTSCHE ABWASSER REINIGUNGS-GESELLSCHAFT (Otto Mohr, inventor) Ger 521,427, Mar 7, 1930

Apparatus for removing acum and sediment separately from waste waters. WILHELM RADERMACHER and CLEMENS DELKESKAMP Ger 503,204, Aug 11, 1925

Furnace (with superposed horizontal conveyors) for drying garbage, etc. JOHN H FEDELER U S 1,797,335, March 24

Septic tank. CECIL H RYMAL (one half to George R. Gregory). U S 1,797,697, March 24

15—SOILS, FERTILIZERS AND AGRICULTURAL POISONS

J I SKINNER AND M S ANDERSON

Recent advances in science: pedology. G W ROBINSON *Science Progress* 25, 595-602(1931), cf C. A. 24, 2822—A review of recent work on soil chemistry.

Chemical-agricultural studies on Italian soils. F ROGAI *Boll 1st super agr* 110 Pisa 6, 493-503(1930)—The chem analyses of 61 soils of Tuscany are given.

Soils of the eastern township of Quebec. R. R. MCKIBBIN AND L. I. PUGSLEY MacDonald College, McGill Univ, *Teck Bull No 6*, 3-64(1930)—All soils in this region have a high percentage of org matter concd in the surface layer. They also have relatively low Ca and P contents and a low nitrate N content throughout the growing seasons. Three main kinds of soil are found: heavily podsolized, brown forest and muck swamp soils. The podsol and brown forest soils have high CaO require-

ments and low pH values. The subsoils of the podsol soils are in most cases better supplied with Ca, Mg and sesquioxide bases than are the surface soils. The forest crops from a large section of this area should be of great value, while the muck swamp tracts should prove highly fertile if they could be drained. J. R. ADAMS

The podsol process in soils. T. J. DUNNEWALD. *J. Am. Soc. Agron.* 23, 259-65 (1931).—An attempt is made to follow some of the details of the podsol process on a lab scale. The results from this method indicate that org. matter is not concerned in the podsol accumulation of SiO_2 but is concerned in the translocation of bases and sesquioxides. Fxtn of the bases and formation of acid conditions appear to be necessary preliminaries to the deposition of Fe_2O_3 and Al_2O_3 in the subsoil. The rapid removal of bases from timbered soil as compared to grass soil is shown, and the slower removal of sesquioxide as compared to the rapid removal of bases is brought out. Available P seems to be lost more rapidly into the subsoil of grass soils than is the case with timbered soils. J. R. ADAMS

Some characteristics of an eroded soil. G. W. MUSGRAVE and HENRY DUNLAVY. *J. Am. Soc. Agron.* 23, 245-52 (1931).—A relatively slight degree of erosion has a marked downward effect upon the yield of cotton. A phys. study of the soil across the area indicated a downward trend of the moisture and of the relative H_2O -holding capacity of the soil loss of org. matter, decline in amt of colloids and moisture equiv. and a lower percentage of clay on the eroded portions of the slope as compared with the uneroded sections. J. R. ADAMS

An extension program for the control of soil erosion in Nebraska. P. H. STEWART. *J. Am. Soc. Agron.* 23, 233-8 (1931).—The use of legume crops to maintain the org. matter content of the soil is emphasized for Nebraska conditions. J. R. ADAMS

A laboratory study of the field percolation rates of soils. C. S. SLATER and H. G. BYERS. U. S. Dept. Agr. Tech. Bull. 232, 1-23 (1931).—A description is given of an auger which was devised to obtain undisturbed cores of soil from the field. The prepn of the sample for the percolation expts. and the exptl. procedure for measuring the rate of percolation are fully described. There is a considerable degree of correlation between the rates of percolation obtained on field cores and on the corresponding lab-packed samples for most of the surface soils studied and a marked lack of correlation for the subsoils. The percentages of silt these soils contain, the suspension percentage and the percolation ratio have been found to represent the same order. The suspension percentage and the percolation ratio are valuable criteria of relative-field permeability. The field-percolation rate of a soil is governed more by the water passage-ways it contains than by the character or vol. of the pore space of the soil mass. The ratios for erosion and for the permeability of soils show that these properties are inversely proportional in terms of the factors of the percolation ratio. J. R. ADAMS

The seasonal changes of some biological factors of forest soils. DÁNIEL FEHÉR. *Matematik és Természettud. Értesítő* 47, 617-41 (in German 642-51). *Math. Naturw. Ber. Ungarn* 37, 51-80 (1930), cf. *C. A.* 23, 3044.—Forest soil contains most humus in September, caused by fallen leaves. The least humus was found in May. Max. humus content and pH are correlated. Their minima do not agree. The humus content of the soil of a black pine forest is greater than that of acacia forests. Black pine should therefore be preferred in reforesting sandy soils in bad condition. S. S. DE FINÁLY

The chemical composition of soils of the Hungarian Great Plain. S. ARANY. *Mezőgazdasági Kutatások* 2, 557-68 (1929).—Thirty g. dry soil is moistened with distd H_2O so as to form a pulp. The amount of H_2O multiplied by 100 and divided by the wt. of soil yields a ratio no. averaging 38-40 for light soils, less than 30 for sandy soils, more than 40 for heavy soils and more than 50 for clay soils. The pH varied between 5.0 and 7.3. Active acidity ($s. e.$, pH detd. in N KCl soln. or titrated as exchange acidity) does not play a great role. The use of slowly decomposing lime fertilizers seems to be preferable. S. S. DE FINÁLY

Data on the composition of soils of the Hungarian Great Plain. II. S. ARANY, G. BUJKE and J. GIDRÓ. *Mezőgazdasági Kutatások* 3, 317-33 (1930).—Plant-food content of soils was examd. Total N was found in considerable amounts. Surface layers contained 66 to 435.5 mg., 1st subsoil (15-55 cm. deep) 46-207 mg., and layers 45-75 resp. (55-85 cm. deep) 20-172 mg. Total P_2O_5 content varied from 55 to 247 mg. P_2O_5 soluble in citric acid moved from 0.6 to 82 mg., K_2O soluble in citric acid between 2-36 mg. $CaCO_3$ content and pH values of the cultivated layers are very varying. The upmost level lacked lime in each case. S. S. DE FINÁLY

Universal apparatus for field investigations of soils. J. DI GLÉRIA. *Kísérlet*

Kölemenek 31, 295 9(1929)—A detailed description is given of a combined Wheatstone bridge and potentiometer. S. S. DE TINKLY

A large sand culture apparatus. FRANK M. EATON *Soil Science* 31, 235-41 (1931)—A sand culture app. is described whereby 50 gals. of nutrient soln. may be used without discarding any portion of it. It is suitable for making cond. tests at frequent intervals in order to follow changes in concn. M. S. ANDERSON

Microanalytical determination of silica in aqueous extracts of soil. A. NÉMÉC, J. LANCÉ AND A. KORPÓCZ *Z. anal. Chem.* 83, 428-45 (1931)—The previously described colorimetric methods for detg. SiO_2 were tested, particularly with respect to manipulative details desirable for SiO_2 in concns. likely to occur in ext. of soil. A method was worked out which is suitable for detg. such quantities of SiO_2 in the presence of as much Fe^{+++} and Fe^{++} as is likely to occur in the soil ext. To the soln. contg. 0.5-0.7 mg. of SiO_2 in 30 cc. add 5 cc. of 1% citric acid soln. and 5 cc. of molybdate reagent (50 g. NH_4 molybdate dissolved in water, treated with 30 cc. of concd. H_2SO_4 and dil. to 1 l.). Shake well and after 5 min. add 7 cc. of 4% H_2SO_4 (by vol.) and 2 cc. of 1% NaVO_3 soln. Then treat with 2 cc. of 5% $\text{Na}_2\text{S}_2\text{O}_4$ soln., and after allowing 15 min. for reduction to take place, add 35 cc. of alk. glycerol soln., (600 cc. of 10% Na_2CO_3 and 100 cc. of glycerol) and dil. to exactly 100 cc. After 2 hrs. compare the blue color with standards in a colorimeter. The presence of Fe, Ca, Mg, Al, As and P does not interfere if the directions are followed carefully. Good results can also be obtained if King's method is modified as follows. To the soln. contg. 0.1-1.0 mg. of SiO_2 in a 100 cc. measuring flask, add 3 cc. of 10% ArOH and 3 cc. of 10% molybdate soln. Shake well, allow to stand 5 min. and then add 2 cc. of a reagent made by dissolving 0.5 g. of hydroquinone in 100 cc. of 20% $\text{Na}_2\text{S}_2\text{O}_4$ soln. After 0.5 hr. dil. with 10 cc. of 10% Na_2CO_3 soln., dil. with water to 100 and compare the color with that of standards. W. T. H.

The relation between buffer power and saturation capacity of the soils. R. NICOLOINI (WITH F. BACCOLI) *Boll. ist. super. agrario Pisa* 6, 207-23 (1930)—The buffer power against the bases of a soil increases with the increase of the absorbing (or satn.) capacity of the colloidal matters contained in the soil, while the buffer power against the acids increases with the increase of the carbonate content of the soil. G. A. B.

Reliability of the quinhydrone method in the determination of pH values of soils. G. HAROS *Mezőgazdasági Kutatók* 3, 189-90 (1930)—The time curve can be shortened by use of a Pt electrode which was previously glowd out and put into dist. water for 1-2 min. Such electrodes show const. potentials in 5 sec. which remains for 10-15 min. Results by this method agree well with those obtained with the Haberklemensiewicz cell. Colorimetric pH detns. often show large divergences.

Hydrogen-ion concentration, aluminum concentration in the soil solution and percentage base saturation as factors affecting plant growth and acid soils. W. H. PEARCE *Soil Science* 31, 183-207 (1931)—A greenhouse study was made of influence of growing crops on pH , Al concn. and degree of satn. of soils formed under different climatic conditions. Each was brought to different degrees of acidity by appropriate lime or acid treatments. The crit. pH for growth of crops varied considerably with different soils. Analyses of displaced soil solns. showed but little Al even in those from the acid soils. The concn. of Al in different soils of like pH varied considerably. There was some correlation between concn. of Al in the displaced soil soln. and growth of corn on acid soils, but there was less correlation between Al concn. and yield of sorghum and barley. No relation was found between the amt. of phosphate in the soil soln. and the concn. of Al at which plant injury developed. The concn. of Ca bore no relation to plant injury. Some correlation was found between the ratio of K to Ca in the soil soln. and plant injury. This ratio varied considerably for different soils at similar pH values. Concn. of Mn showed no relation to plant injury from soil acidity. Soils varied considerably in their percentage base satn. at similar pH values. Good correlation was found between percentage base satn. of acid soils and plant injury. The proportions of the various bases present in the exchange complex are thought to be a factor affecting plant growth on acid soils. M. S. ANDERSON

Temperature effects on the exchange acidity of lowland and swamp soils, especially the effect on the amount of acid ferrous iron. K. KAWAMURA *J. Sci. Agr. Soc. (Japan)* 1928, 1-20 (in English 21)—Acid soils were treated with normal KCl soln. at varying temps. (4°, 16°, 26° and 36°), and the titratable acidities and the amts. of Al and Fe in the exts. were detd. quantitatively. The acidity found agreed very well in every case with that calcd. from the amts. of Al and Fe, an indication that the acidity was produced by base exchange. The lowland and swamp soils, (peat soils and rice-

field soils), which contained more or less org. matter, produced a considerable amt. of FeCl_3 in the KCl soln., and the amt. increased at a marked rate as the treating temp. rose. Org. (arol) soil, however, did not liberate any Fe in the soln. In most cases with all soils, the amt. of AlCl_3 produced decreased to a certain extent with the rise of temp. At the same time the sought titratable acidity decreased also, because the AlCl_3 generally played a greater part in the acidity than did the FeCl_3 . The influence of the treating temp. is explained as follows: (1) As shown in the expt. with Al^{+++} and Fe^{+++} humates pregel from peat and cane sugar, the base exchange of the ferrous humate is markedly accelerated by rise of temp., whereas that of Al humate is somewhat hindered. (2) The higher the treating temp. the greater the amt. of FeCl_3 produced by the side reaction between AlCl_3 formed by base exchange and FeO originally present in the soil. This part of the AlCl_3 is changed to an insol. form, as illustrated by the reactions between the soil and old AlCl_3 solns. E. SHIMIZU

Practical methods of alkali soil reclamation applied on the Hungarian Great Plain, S. ARANY *Mezőgazdasági Kutatók* 4, 11 231(1931). A general description is given of 3 methods: (1) liming, (2) covering the surface with a calcareous marly subsoil and (3) covering and mixing the surface layers with drift sand. S. S. DE FÉVÁLY

Data on the liming of Hungarian soils and on the magnesium content of alkali soils, S. ARANY *Mezőgazdasági Kutatók* 2, 1 11(1929). Sandy soils of Nyíregyháza showed neutral to weakly alk. reactions and some hydrolytic acidity. Careful liming is proposed and eventually the use of Thomas slag or Rhénania phosphate. Well water, soil water and most of the whole profile of alkali soils contained Mg . The quantity of water sol. Mg varies: that of Mg sol. in salt solns. is the same as in Californian alkali soils. S. S. DE FÉVÁLY

Determination of small amounts of calcium carbonate in soils, L. DWORÁK *Kísérlet Közlönyek* 31, 174 81(1928). Ten and 20 g. of soils contg. no Na_2CO_3 are treated with N NH_4Cl soln. and titrated with 0.1 N HCl . Percentage CaCO_3 equals $(2A/B)(B-A) + [(2A/B)-1](B-A) \times 0.005 \times 10$, where A means cc. of 0.1 N HCl used for titrations of 250 cc. soln. of 10 g. soil and B the same for 20 g. soil. The method gives good results. S. S. DE FÉVÁLY

Reclamation experiments on Hungarian soils rich in lime and containing sodium carbonate, J. DE GÁFRA *Kísérlet Közlönyek* 32, 252 89(1929).—The pH and phys. condition of soil bear a relation to the Ca in concn. of the soil soln. Calcium coagulate perfectly only in soils with pH under 7.0. Coagulation by Ca salts was observed when the soil soln. contained at least 0.278 mg. equiv. Ca after treatment. Equiv. quantities of reclaiming agents do not show the same results. The least quantity is needed of H_2SO_4 and HCl , somewhat more of AlCl_3 , $\text{Al}_2(\text{SO}_4)_3$, and yet more of Ca salts. Reclaiming agents should be completely washed out of reclaimed soil. CaCO_3 was ineffective with these soils. S. S. DE FÉVÁLY

Problems of heavy metals in soil chemistry, LÁSZLÓ DE PUTNOKY, *Mezőgazdasági Kutatók* 4, 60 65(1931).—The role of heavy metals in the animal products, soils and fertilizers is discussed. P. calls attention to the necessity of regulating the heavy metal content of the soils, especially that of sandy soils. Z. I. K.

Comparison of the Dirks and Scheffer and the Mitscherlich methods for determining the phosphoric acid requirements of the soil, LAMMERN, *Das Superphosphat* 7, 61-2(1931).—In 225 comparative expts. with soils ranging from rich to very poor in P_2O_5 the 2 methods gave results that agreed in 222 cases. The poorest agreement was obtained on soils that were not deficient in P_2O_5 . K. D. JACON

The significance of colorimetric methods for the determination of phosphoric acid in soils, L. FRUM, *Das Superphosphat* 7, 67-8(1931).—The Molybdenum colorimetric and the Lorenz gravimetric methods were compared for the detn. of P_2O_5 in citric acid exts. of soils. The 2 methods were in good agreement when the P_2O_5 content did not exceed about 60 mg. per 100 g. of soil. With larger amts. of P_2O_5 the colorimetric method gives results that are too low. The presence of citric acid causes low results by the colorimetric method if the aliquot for analysis contains more than the equiv. of about 10 cc. of a 1% soln. of citric acid. If carried out under carefully controlled conditions, the colorimetric method gives good results, and because of the speed with which it may be carried out it is very desirable for use in the citric acid method for the routine detn. of the P_2O_5 requirements of soils. K. D. JACON

Changes in the availability of phosphorus in irrigated rice soils, R. P. BARTHOLOMEW, *Soil Science* 31, 209-17(1931).—Studies were made under field conditions of the effect of irrigation upon pH and available phosphates in rice soils. The irrigation water contained considerable $\text{Ca}(\text{HCO}_3)_2$ and rendered the surface soils more alk. than the subsoils. When phosphate fertilizers were used, the irrigation water caused a decrease

in water sol inorg phosphate. This action was due partially to reversion of the phosphates added. Three months after irrigation, org water-sol P had increased, probably because of the activities of bacteria. Phosphatic fertilizers are not recommended for rice in this region when calcareous water is used for irrigation. M. S. ANDERSON.

Effect of carbon on vegetation. III. R. PEROTTI AND C. FERRETTI. *Boll. ist. super. agrario Pisa* 6, 147-48 (1930), cf. C. A. 23, 5531.—Both animal and vegetable C are effective in rendering harmless the toxins produced by microorganisms and in activating the respiratory process. A new apparatus is described for the detn. of the respiratory activity of the soils. The action of the C continues for one year after it has been added. G. A. BRAVO.

Effect of carbon on tobacco culture. G. BONUCELLI AND V. CINTI. *Boll. ist. super. agrario Pisa* 6, 339-54 (1930).—The expts. described in the preceding abstr. were made on tobacco. If powd. vegetable C is added to the soil, the crop is increased, and the leaves show greater burning capacity. G. A. BRAVO.

A total-carbon procedure [for soils]. LAUC WINTERS AND D. C. WINTER. *J. Am. Soc. Agron.* 23, 290-5 (1931).—This is a modification of the Winters and Smith method (cf. C. A. 23, 5531) for the detn. of total C, in which the loaded boat can be moved from the cool part of the tube to the center of the furnace after all the connections have been made and the O flow started, the active catalyst is in large enough quantity to insure oxidation even under extremely rapid flow of combustible gas. The loaded boat is placed in that portion of the tube outside the furnace with an Fe "pusher" behind it. The connections are all made and the O flow is started. A solenoid which fits over the tube is then centered over the Fe "pusher," the circuit made and the boat is gradually pushed into the furnace. This method of inserting the charge reduces the possibility of gas losses. The percentage error in this method is $\pm 0.55\%$. J. R. ADAMS.

The transformation of citric acid into acetone by the soil. R. NUCCORINI (with A. FELICIANI). *Boll. ist. super. agrario Pisa* 6, 241-8 (1930).—The decompn. of citric acid and the formation of acetone by the soil take place in the presence of ferric salts and Fe-contg. fertilizers (Thomas slag). This decompn. is attributed to the presence of iron compds. and also of other, unidentified compds. (1-3%). Cf. Dean and Dean (C. A. 24, 3588). G. A. BRAVO.

A study of some soils for cotton from the Segou region made in the laboratory of the Alsace Potash Co., J. FRANCK DE FERRIERE AND E. NATIER. *Rev. botan. appl. agr. trop.* 11, 84-98 (1931).—These soils along the Niger River vary from the coarse sand to the heavy clay type. In order to obtain good yields all these soils require the addn. of fertilizers high in P_2O_5 and K_2O . The fertilizers should also contain N, when they are to be used on the sandy soils. The geological formation of these soils is discussed on the basis of their phys. constitution. J. R. ADAMS.

The effect of the soil reaction on beet yields. JORRÉ. *Compt. rend. acad. agr. France* 17, 123-32 (1931).—In the heavy plateau soils the optimum soil reaction for beets should be between pH 7.2 and 7.5. On slightly acid soils the max. yield is sometimes lowered as much as 10%, and this depressive effect can be increased by the addn. of physiologically acid fertilizers. The depressive effect of the acid reaction on the yield is somewhat corrected by the addn. of $Ca(NO_3)_2$ to the soil. J. R. ADAMS.

Determination of nutrient requirement of soils by means of electro dialysis. J. DE GLÉRIA. *Mémoires de l'Institut de Chimie* 3, 185-8 (1930).— P_2O_5 and K data detd. by electro dialysis are always higher than those detd. according to Neubauer. Neubauer data did not show any regular connections with dialysis data. Also NO_3 and NO_2 contents of soil can be detd. by electro dialysis. S. S. DE FIVELY.

The supply of plant nutrients in the soil determined by Neubauer's method and calculated by the static or Mitscherlich methods. CLAUSEN. *Z. Pflanzenernähr., Düngung u. Bodenk.* B10, 145-55 (1931).—Available P and K in soil of plots used for fertilizer expts. for 8 years previously were extd. by the ryse-seedling method of Neubauer. The indications are compared with results to be expected from the known addns. in fertilizer and removal in crops (static), and also in some instances with results obtained by growing crops to maturity in pots (Mitscherlich). Available P by the Neubauer test was much less than expected except in plots fertilized without P. Mitscherlich indications for P were usually in good agreement with the calcd. P supply except on unfertilized soil. The Neubauer procedure seems to be a better indicator for available K than that of Mitscherlich, which always furnished lowest indications except on soils fertilized without K. C. J. SCHOLLENBERGER.

The "Pehameter" and its practical usefulness. R. HERRMAN. *Z. Pflanzenernähr., Düngung u. Bodenk.* B10, 142-5 (1931).—The "Pehameter" is an app. intended for the

dett. of soil pH , it consists of a porcelain plate with a depression at one side, from which a narrow channel extends across the face. On each side are small rectangles painted with ceramic colors red to blue, which indicate characteristic tints with the "Universal" indicator from pH 4 to 9 by single units. A little soil is placed in the cavity, said with indicator, and after a min a drop is led down the channel and its color compared with the standard. Tests with 30 soils of known pH , as detd. by standard methods, gave approx. correct indications in 11 cases, an error of 1 pH or less in a total of 23 cases and over 2 pH in 3 cases. The largest errors were with moderately acid soils, pH 4.5–5.5, which were indicated to be neutral. C. J. SCHOLFFENBERGER

New principles in determination of fertilizer requirement of soils. Importance of relative examinations. L. DWORAK *Mezogazdasági Kutatások* 3, 355–9 (1930).—Not only requirements but also possible and probable influence of fertilizers should be examined by simultaneous expts. on untreated soils. Also exptl. failures are removed if relative data of unfertilized and fertilized soil are detd. S. S. DE FINÁLY

Determination of fertilizer requirements of soils. J. BECKER *Mezogazdasági Kutatások* 1, 65–91 (1928).—Data on P_2O_5 and on K requirements obtained by means of the methods of Lemmermann, König, Hasenbaumer, Sigmond, Niklas and Neubauer agreed. The citrate and Neubauer methods gave too low results on easily sol. P_2O_5 in soils contg. much carbonates. Results by the method of Sigmond on total P_2O_5 were high on some soils. K requirement was detd. by a modification of Sigmond's method and agreed with results by the Neubauer method. S. S. DE FINÁLY

New method of illustrating the results of soil analysis. Z. I. KERTÉSZ AND J. CSIKY *Mezogazdasági Kutatások* 2, 201–8 (1929).—Data are shown as differently colored sectors of a circle. Characteristics of soils can be seen at a glance. S. S. DE FINÁLY

The soil microbiology of the lower Val di Cecina. O. VERONA AND A. DEL TREDICI *Boll. ist. super. agrario Pisa* 6, 71–102 (1930).—The chem. analyses of the soils are reported. Two new species of bacteria were found *B. danielis* and *B. flavum non-liquefaciens*. The seasonal bacterial activity, in regard to N circulation, is studied; this activity is defined by (1) ammonification power (mg. of NH_3 produced from 1 l. of soil), which is max. in summer, (2) nitrification power (mg. HNO_3 produced), (3) denitrification power (time, in hrs., necessary for the disappearance of nitrates in the soils), which is very low in winter, (4) N-fixation, which is greater in spring and lower in the subsoil. G. A. BRAVO

The microbiological activity of Hungarian alkali soils. D. DANIEL FERÉZ AND REISZ BOKOR *Motemohák és Természettudományi Értesítő* 47, 270–319 (in German 320–36) (1930).—The alkali soil of Püspökladány has a special microflora with many ray-like organisms. Soil should also be reclaimed biologically. Liming influences quantitatively the compn. of the microflora. Young forests should be helped by conserving the moisture content of the soil on reforesting. The sp. microflora of stall manure cannot live in alkali soils. Aerobic N-fixing microorganisms generally are absent in alkali soils. Decompn. of cellulose goes on very slowly. Respiration and CO_2 production are almost normal. Alkali soils contain little N, little nitrite and nitrate. Soils contg. *Achillea* assoc. with *Inula britannica* contain most nitrate N, nursery soils contained very little nitrate N. The content in N decreases in the lower layers, and also the no. of nitrifying bacteria. Nitrification was much faster in soils contg. 0.2% Na_2CO_3 on the addn. of $(NH_4)_2SO_4$ or on inoculation with bacteria. The ammonifying capacity of alkali soils having no Na_2CO_3 or contg. less than 0.03% total salt is excellent. S. S. DE FINÁLY

Investigations on the protozoal fauna of forest soils. D. DANIEL FERÉZ AND LAJOS VARGA *Matematik és Természettudományi Értesítő* 46, 235–71 (in German 272–6) (1929), *Math. Naturw. Ber. Ungarn* 37, 81–110 (1930).—The protozoa consist for the most part of amoebae. Two maxima occur in the periodic changes of protozoa numbers, one in Nov. and Dec. and the other in the beginning of summer. The maxima are probably caused by soil humidity. No definite correlations can be observed between the life activities of soil bacteria and soil protozoa. No large difference was found in the protozoal fauna of conifer forests and that of leafy forests. Forest soils seem to contain fewer protozoa than do soils of gardens and fields. S. S. DE FINÁLY

Experiments on the biological activity of soil protozoa. L. DE TELEGDY KOVÁTS *Kísérlet. Közlemények* 31, 223–31 (1928).—Daily changes in the numbers of amoebae and bacteria in good garden soil were investigated. The data check for Hungarian soils the theory of Russel and Hutchinson. Protozoa seem to have a stimulating effect on *Azotobacter* cultures. *Azotobacter* cultures vitated with *Colpidium* and *Paramecium* fixed more N than pure ones. S. S. DE FINÁLY

Electrophoretic measurements of the root-nodule bacteria of Leguminosae. F.

ZUCKER *Mezőgazdasági Kutatószok* 3, 49-64(1920)—Representatives of the alfalfa, red clover, pea, bean, soy bean, cowpea and lupine groups were studied. The effect of the degree of dilution of cell suspension was determined from the electrophoretic velocity. Velocities of dilute solutions are higher than those of concentrated solutions. Velocity is influenced by the amount of lime and acid by products of organisms. Variations of velocities of different strains of groups of bacteria never exceed 2 microns per sec. if 22.5 v. is used. Various strains of *B. radiobacter* showed different velocities which average about the same as those of nodule bacteria. S. S. DE FIVELY

The influence of combined nitrogen on growth and nitrogen fixation by *Azotobacter*. JAMES L. FULLER AND LEO J. REITIGER *Soil Science* 31, 219-24(1931)—*Azotobacter* is able to fix substantial quantities of free atm. N when cultivated in a medium free from fixed N and in an atm. free from NH_3 and HNO_3 . A study is made of the influence of a large no. of nitrogenous org. compns. upon the fixation of N. In general the more complex org. compns. except peptone, are not actively attacked by *Azotobacter* while the simpler ones including the lower amino acids are readily utilized. Where an abundance of combined N in suitable form is present, very little atm. N is fixed. M. S. ANDERSON

The germination of wheat in alkaline and in acid media. Gy. ERŐJESSY *Mezőgazdasági Kutatószok* 1, 93-111(1924)—Wheat types grown on alk. soils endure NaOH and Na_2CO_3 more readily than do those accustomed to neutral or acid soils. The latter endure acids more readily. Small amounts are stimulating. NaHCO_3 is more toxic than Na_2CO_3 in the same concn. H_2SO_4 is more toxic than H_3PO_4 . The toxic action of Na_2CO_3 can be decreased with CaSO_4 and humus added. S. S. DE FIVELY

Studies on root growth of wheat types accustomed to alkaline and acid soils. Determination of osmotic pressures of wheat. Gy. ERŐJESSY *Mezőgazdasági Kutatószok* 2, 481-93(1929)—Root growth is better in Na_2CO_3 or NaHCO_3 solutions if wheat types accustomed to alk. soils are used. Wheat accustomed to acid soils prefers acid media. The osmotic pressure of improved wheat equalled 30.5, that of unimproved wheat 29.5. S. S. DE FIVELY

Yields of crude and assimilable nutrients obtained from meadows mowed twice and three times, after various fertilizer treatments. H. JANSON AND W. KIRSCH *Illus. Arch. Landw., Alt. B., Tierernähr. Tierzücht.* 1, 166-74(1930)—Plots of meadow land were fertilized with N, K and P, alone, and combinations of all 3, at the beginning of the expt. and after each cutting. Total yields and yields of all constituents except crude protein were larger from the plots mowed twice during the summer, although there was no appreciable difference in the yields of assimilable nutrients, based on a starch basis. W. CORDON ROSE

Effects of various plant foods on growth activities and development of oats. C. K. McCLELLAND *J. Am. Soc. Agron.* 23, 304-11(1931)—Under the given exptl. conditions P generally seemed most beneficial in increasing the no. of tillers, the no. of heads at harvest, the no. of spikelets per head, the yield and the size of the seed of oats. It also exerted a marked influence in causing early tillering and growth and a quick shading of the ground. N acted very similar to P in its effect upon tillering, no. of heads at harvest, no. of spikelets per head and yield. K exerted no influence, unless one of restriction, on no. of tillers and no. of heads per plant at harvest and upon yield. Combined with N and P, it increased the size of seed, but the latter elements exerted better influence on the wts. per bushel and upon general growth and development. J. R. A.

The study of the black cotton soil under different fertilizer treatments. R. D. REGE *Proc. 15th Indian Sci. Cong.* 1928, 50-1—Plot culture experiments with jowar were conducted under various treatments. Conclusions: (1) The Rab process (burning of cow dung on the plot) is not suitable to the tracts with low rainfall. It leads to rapid increase in the nitrification, but the crop growth is not correspondingly increased. This treatment also decreases the colloid content of the soil, especially in the first 4 inches, but this result is only short lived. (2) Safflower cake produces the best effect. There is not only vigorous growth of the plants, but this treatment leads also to early flowering and seeding. Water logging of these plots for 2-3 days does not produce any bad effect on the crop. (3) Wt. for wt., the effect of karanj cake is less than half that of the safflower cake. The N content of the karanj is also about half that of the safflower, and the difference in the plant growth may be attributed to this. (4) $(\text{NH}_4)_2\text{SO}_4$ produces as good an effect as safflower if buffered well with lime. (5) The bad effect of water logging of the control soil is not found to be due to either increased CO_2 content in the soil or decreased soil aeration. The nitrification in the soil is very little. E. J. C.

Fertilizing with carbon dioxide. Gy. ERŐJESSY *Természettud. Közöny* 62,

630-5(1930)—Expts carried out in Germany by Reinau are described

S S DE FINÁLY

Potassium and phosphorus fertilizer experiments with barley in Hungary during 1929. J SURÁNYI *Kisérlet Közlemények* 33, 402-12(1930)—Use of both gave good results on medium heavy soils

S S DE FINÁLY

Experiments with potassium-phosphoric acid fertilizers on beet barley in 1928. J SURÁNYI *Kisérlet Közlemények* 32, 239-50(1929)—The use of both K and P_2O_5 fertilizers gave good results on Hungarian barleys. Not only the quantity but also the quality of the barley crop was better

S S DE FINÁLY

Phosphate-containing artificial fertilizers with respect to their raw materials. L. SOLT *Chem Rundschau Mitteleuropa u Balkan* 6, No 14, 68-9, 76-80(1929)—A general review

S S DE FINÁLY

Behavior of Rhenania phosphate in the soil. L. DWORAK *Kisérlet Közlemények* 32, 430-42(1929)—The soly of Rhenania phosphate added to loam and to sandy soils did not show any decrease in 5 months

S S DE FINÁLY

Determination of phosphoric acid content of Rhenania phosphate. S TANGI BOGSCHE *Kisérlet Közlemények* 33, 200-2(1930) P_2O_5 is dissolved in Petermann's soln, and phosphoric acid is detd with ammonium molybdate according to Lorenz

S S DE FINÁLY

Superphosphate enriched with ammonia. S I VOLFKOVICH, L E BERLIN, I L HOFMAN AND A A IONAS *Udobrenie i Urochas (Fertilizers and Crops)* 2, 556-69(1930)—Two kinds of raw phosphate were tried out in the prepn of ammoniated superphosphate by the use of H_2SO_4 and the phosphates of ammonia. From Portland phosphate with mono ammonium phosphate and H_2SO_4 a product was obtained contg 21% P_2O_5 , with 10.5% available, and 10% N. The amt of H_2SO_4 used was 15% above that required by the stoichiometric equations, and its strength was diluted to 40% by the use of a 26% soln of $NH_4H_2PO_4$. When a mixt of mono and di ammonium phosphates was used, the product contained 22% P_2O_5 , with 18% available, and 3% N. The H_2SO_4 was also 15% above the stoichiometric and was diluted to 35% with 26% $NH_4H_2PO_4$. With $(NH_4)_2HPO_4$ alone the product contained 23% P_2O_5 , with 19% available, and 5% N. The H_2SO_4 was 10% above stoichiometric, and it was diluted to 25% with a 26% soln $(NH_4)_2HPO_4$. For the Saratov phosphates triammonium phosphate (in liquid form) with H_2SO_4 was used, and the product obtained was 22-23% P_2O_5 , with 20.5% available, and 10% N. The optimum conditions were 100 parts (by weight) of phosphate for 80 parts of $(NH_4)_2HPO_4$, for each unit weight of $(NH_4)_2HPO_4$ 0.5 part of 60 Bé H_2SO_4 was used besides what is necessary to convert the raw phosphate into superphosphate. By using a satd soln of $(NH_4)_2HPO_4$, the product contained 18% P_2O_5 , with 17% available, and 5% N. The reagents used, were 100 parts of phosphate, 100 parts of the satd soln with 0.74 part of 60 Bé H_2SO_4 for each part of $(NH_4)_2HPO_4$, besides what is necessary for the production of ordinary superphosphate. The best product from the standpoint of phys conditions was obtained by the use of the mixt. of mono- and di ammonium phosphate. Production on a semi factory scale is under way with the mixts mentioned

J S JOFFE

Some fermentation characteristics of various strains of *Rhizobium meliloti* and *Rhizobium japonicum*. R H WALKER AND P E BROWN. *Soil Science* 30, 219-29(1930)—A study was made of the fermentative characteristics of 23 strains of *Rhizobium japonicum*. These organisms were grown in yeast water glucose and yeast water-galactose media. The change in H ion concn was taken as a measure of fermentative ability. The fermentative ability of a particular strain is comparatively const. Different strains, however, show wide differences in fermentative ability. Some strains of *meliloti* produced a distinctly alk. reaction in the media, while others produced considerable acidity in media of the same compn. The *japonicum* strains, as a rule, produced more alk. than did the *meliloti* strains. The *meliloti* strains fermented glucose and galactose with about the same results, while the *japonicum* showed more differences in this respect. The results indicate that fermentation tests under the condition followed would not serve to give a distinct sepn. of these 2 species of legume bacteria

M S. ANDERSON

Fermentations in a heterogeneous and discontinuous medium. A. DEMOLON AND G BARBIER. *Compt rend* 192, 514-5(1931)—In a sandy soil diffusion and migration are easy, but in a medium contg siliceous clay these phenomena are depressed

T. H RIMER

A comparison of some nodule-forming and non-nodule-forming legumes for green manuring. LEWIS T. LEONARD AND H R. REED. *Soil Science* 30, 231-6(1930)—Legumes not bearing nodules are compared with legumes bearing nodules in green

manuring expts. The indicator crops grown after plowing under the legumes show no decided difference in these 2 types of legumes as affecting total N, dry wt. and height. M. S. ANDERSON

Amount of manure necessary for vegetable growing. F. K. CRANDALL AND T. E. ODLAND. Rhode Island Agr. Expt. Sta., *Bull.* 225, 5-31 (1930).—Field expts. conducted for 5 years included early beets, spinach, carrots, cauliflower and eggplants fertilized with various combinations of stable manure, green manure crops and com. fertilizers. Eggplants and cauliflower showed the greatest response to N, carrots were little affected. Spinach and cauliflower showed the greatest response to P. Manure contg. shavings was practically as effective as manure contg. straw. Manure increased crop yields when even moderate treatments of com. fertilizers were used. It is possible to substitute green manure and com. fertilizer for a considerable part of the stable manure which is ordinarily applied to vegetable crops. C. R. FELLERS

Nitrogen loss and action of cold and hot manures. H. D. MEYER, P. OBST, F. WILCZEWSKI AND W. DIETRICH. *Z. Pflanzenernähr., Düngung u. Bodenk.* 10B, 121-36 (1931). cf. *Deut. landw. Presse* 56, 629-43 (1929).—Cow manure solidly packed in tight pits was little changed in appearance and had a slight NH_3 odor when removed 4 months later. Loosely packed manure was well rotted and without NH_3 odor. The temp. rose to 30° in the former case and to 55° in the latter case in 6 days. Losses in dry matter and total N were 10.1 and 7.5% for cold and 9.1 and 15.1% for hot fermentation, resp. Pot. and plot expts. did not furnish any conclusive evidence of superiority for manure stored by either method. The relative value of different methods for storing manure can be detd. only by long-continued field expts. on a considerable scale. C. J. SCHOLLENBERGER

The chemistry of artificial manure making and materials used. F. ZUCKER AND L. DE TELEGDY KOVÁTS. *Mészgazdasági Kutatóköt.* 3, 193-206 (1930).—Substances giving a slightly alk. reaction should be applied. The conditions prescribed by Krantz for "Edelmist" manuf. should be followed in order to obtain a well rotted stall manure. S. S. DE FINÁLY

New experimental data on the composition of straw manures. L. DE TELEGDY KOVÁTS. *Mészgazdasági Kutatóköt.* 3, 243-50 (1930).—The feasibility of artificial manure production was proved by expts. on a larger scale. The principles of Krantz should be followed. Loss of N was 21%, that of P_2O_5 14.7% and that of K_2O 1.7%. Artificial straw manure contained 44.17% N, 29-0.5% P_2O_5 and 21-0.6% K_2O . S. S. DE FINÁLY

Artificial stall manure. J. BOCSE. *Kiserlet Közlemények* 33, 1-15 (1930).—Production of fertilizer from straw, comstalks, leaves, etc. should be undertaken at first by aerobic fermentation, and later air must be shut off. Natural manure should be mixed with the lower layers or natural manure juice should be sprayed on to inoculate the matter with bacteria. S. S. DE FINÁLY

Manufacture of artificial stall manure from straw. L. DE TELEGDY KOVÁTS. *Kiserlet Közlemények* 31, 347-72 (1928).—Straw was fermented under aerobic conditions with addn. of (1) Adco powder, (2) a mixt. of nutrient salts and (3) P compds. and dung water. The product in each case was similar to natural manure. Anaerobic fermentation was not so effective. As for biol. factors, multiplication of thermophilic bacteria is the first stage, then follow molds and later bacteria. S. S. DE FINÁLY

Manufacture of artificial manure from peat. L. KOTEMANN. *Mészgazdasági Kutatóköt.* 3, 324-8 (1930).—The acid reaction of peat was neutralized with CaCO_3 and the mixt. was inoculated with natural stall manure. CaCN_2 was added as nutrient, in quantity corresponding to 0.7% of the org. dry matter content. The product after 4 months had better action in field expts. than did natural manure. Loss of N amounted to 12%, that of P_2O_5 3.1% and that of K_2O 2.1%. S. S. DE FINÁLY

Practical method of manufacturing "Edelmist". H. CLATHE. *Mészgazdasági Kutatóköt.* 3, 376-9 (1930).—Detailed prescription of method. S. S. DE FINÁLY

Spreading of "Edelmist" manufacture in Germany and observations on the quality of the product. H. CLATHE. *Mészgazdasági Kutatóköt.* 3, 463-9 (1930).—Edelmist should be in piles for at least 3 months. It contains more total N than does natural manure, and its NH_3 content is also higher than its content of water-sol. N. Frequent moistening of completed stocks of Edelmist proved to be of great value. S. S. DE FINÁLY

Complete weed-eradicating fertilizers. JEAN DETERRE. *Engrais* 46, 47-9 (1931).—A review is given of the work of Jaguenaud (C. A. 25, 1322) on complete weed-eradicating fertilizers, their use and advantages are stressed. J. R. ADAMS

A review of methods for the evaluation of pyrethrum. C B GNADINGER *Soap* 7, 97-101, 117(1931)

The use of coal-tar products in spraying fruit trees. I' PROFFT *Zentr. Bakt. Parasitenk.* Abt. 11, 83, 127-64(1931)

Pear mealy bugs and results of experiments for their control. F W PETTEY AND C J JOUBERT *Union S Africa Dept Agr. Sci Bull* 95, 23 pp (1930), cf. C. A. 24, 5417 —Oil sprays give the best control of pear mealy bugs (*Pseudococcus maritimus* and *P. gahani*). No more than 1 application of the usual winter spray of a diln contg about 4% actual oil of specifications ranging from 50 to 70% unsulfonatable residue and about 110 sec Saybolt viscosity can be applied with safety to pear trees in S African conditions. Modern proprietary winter oil (petroleum oil) sprays generally possess these specifications. No more than 2 or 3 summer oil sprays of a diln contg 0.8% actual oil with a max of 65 to 70 sec Saybolt viscosity and a min of 85% unsulfonatable residue should be applied, with or without nicotine sulfate or Pb arsenate, to pear trees in foliage until repeated tests over a no. of years prove that the trees will tolerate more.

Experiments with new methods for the control of codling moth in western districts of the Cape Province. I W PETTEY AND M C MOSSOP *Union S Africa Dept Agr. Sci Bull* 96, 50 pp (1930). cf. C. A. 24, 2228, 3500 —The Reinette du Canada and Newtown varieties of apples are particularly susceptible to injury from repeated applications of 1% medium or light summer oil sprays. Addn of summer oil to Pb arsenate sprays, particularly the heavier grades, retards somewhat the removal of the As spray residue, but this difficulty is minimized if Ca caseinate spreader is used and the oil-arsenate spray is applied at least 6 weeks before the fruit is harvested. An artificial cryolite, used at the rate of 2 lb with 0.25 pint of fish oil in 40 gallons of water, applied through the season at intervals of 3 weeks, gave better results in codling moth control than a similar normal program of Pb arsenate only and seemed to have considerable effect in controlling mealy bugs on Kieffer pears. Certain samples of cryolite caused severe foliage burning, while others did little damage. Cu cyanide, 1 lb in 40 gallons, controlled the codling moth as well as Pb arsenate, but caused russetting of the pears. A 7-10% soln of crude molasses, with or without the addn of yeast, was a satisfactory bait for use in codling moth traps. Extensive data on the results of the investigations are tabulated.

A preliminary note on the successful treatment of the jasmine bug (*Antestia cruciata*) by dusting calcium cyanide. RAMACHANDRA RAO *Proc 15th Indian Sci Cong* 1928, 51-2 —Dusting with $\text{Ca}(\text{CN})_2$ was found to have the effect of stupefying the bugs on the bushes and bringing them to the ground. As a certain proportion of the bugs generally revived within a short time, these had to be hand-picked and destroyed by dropping them in kerosenated water.

The onion maggot (*Hylemyia antiqua*) in Ohio, 1929. M P JONES *Econ Entomol* 23, 394-8(1930) —This is a demonstration control project with Bordeaux-petroleum oil emulsion and com oil emulsions against the onion maggot. About 45% increase in onion yield due to the spray is claimed. An ingenious combination power sprayer and cultivator for onions is described.

Ecology of the microbe causing brown spots on tobacco leaves. A GULYÁS *Kísérlet. Közlemények* 33, 279-302(1930) —Spots are caused by an organism called *Alternaria brassicae* var. *tabaci* by Preisser. Methods of prevention consist of treating seeds with a 0.2% soln of Germisan or a 0.25% soln of Uspulum. Plants in the warm beds should be sprayed weekly with 0.5-1% Bordeaux soln, and soil of warm beds must be disinfected with Cl_2O_3 3 weeks before sowing. Ten to 12.5% Cl_2O_3 soln should be used on 1 sq m area. Soil should contain as much lime as required to show alk reaction.

Beet infections with *Cercospora beticola* Sacc. A CIRZANOWSKI *Spisek wydany nakładem Rady Naczelnej Polskiego Przemysłu Cukrowniczego* 1927; *Liisy Cukrowar* 49, No 22, Rozhledy 14 —*Cercospora beticola* occur sporadically decennially. Early beets were more subject to infection than later ones. Plots of beets on which legumes were harvested previously were highly resistant to infection because of the high porosity and N content of the soil. Fertilizing the beets with 0.3 of the annual NaNO_3 fertilizer during July, and fertilizing with P increased the resistance of the beets to infection. Artificially infected or poorly developed seeds showed a greater incidence of infection in the field growths than did the controls. C recommends soaking the beet seeds in dil HCHO before planting.

Influence of seed-preserving agents on the germinating activity. ÁRPÁD SZÜCS *Mézőgazdasági Kutatások* 4, 24-8(1931) — CuSO_4 had a very bad influence on germi

nation activity. Lime added to CuSO_4 decreased this influence. CuSO_4 heated to dust and used as a powder had almost no bad influences on germination. Its seed preserving action was quite good. S S DE FINELY

Increase of action of copper salt solutions used as seed preservatives. S TERÉNYI (*Chem. Rundschau Mitteleuropa u. Balkan* 116, 8, 1-3(1931))—The effective dosage is diminished to 25-30% of original value if 11g salts are added to Cu salts, e.g., 0.25% $\text{Cu}(\text{NO}_3)_2$ and 0.005% 11gBr₂. S S DE FINELY

Laboratory examination of "Germisan-Kurzbeizverfahren." SANDOR TERÉNYI. *Kisérlet. Kótelevények* 31, 332-5(1928)—The germinating power of wheat was not influenced by 0.5, 0.75, 1 and 1.5% solns. of Germisan in the ratio of 6, 4, 3 and 2 l to 100 kg. wheat. Germisan killed the spores of wheat not very strongly infected (not more than 0.1% spores) quite well. Field expts. are under way. S S DE F

Investigations on the seed-preserving influence of copper sulfate-mercuric chloride mixtures. D RÉVY AND I LUDWIG. *Kisérlet. Kótelevények* 32, 299-304(1929)—The addn. of HgCl_2 (equiv. to 0.3-2.4% Hg) increased the influence of CuSO_4 in lengthening the period during which germination is possible. The insecticidal and germicidal action will be investigated later. S S DE FINELY

Sodium chlorate as a herbicide. ALFRED A HANSEN. *Proc. Indiana Acad. Sci.* 38, 139-42(1929)— NaClO_3 had previously been used to kill Canada thistle, field bindweed (*Convolvulus arvensis*) and quack grass. The NaClO_3 is successfully used as a herbicide for poison ivy, buck brush (*Symphoricarpos symphoricarpos* L.) and purple barberry (*Berberis vulgaris* L.). Salt is applied in soln. contg. 1 lb. per gallon. The danger from fire is not great if proper precautions are taken. H M STARK

Waste gas-cleaning material as weed destroyer. ANON. *Chem.-Wig* 55, 190-1(1931)—Waste NaClO_3 from dye works has been used as a weed destroyer but is inconvenient because it must be spread at least twice a year. The waste material from gas-cleaning plants contains a thiocyanate, which dissolves in rain water and penetrates the roots of the weeds. The present price of \$3.75 per ton is prohibitive, however. CURTIS L. WILSON

Prickly pear and its eradication. C R VAN DER MERWE. *Union S. Africa. Dept. Agr., Sci. Bull.* 93, 32 pp(1931)—Spraying with a soln. of 3 lb. As_2O_3 in 1 gallon of water kills all the young growth on old prickly pear trees and also trees up to approx. 3 years old. Spraying with a soln. of 2 lb. As_2O_3 in 1 gallon of water kills all of the above ground portion of the jointed cactus (*Opuntia aurantiaca*) but does not kill the underground bulbs or tubers. Addn. of NaCl to the sprays apparently has no effect on the results. K. D. JACOB

Determination of H ion concentration in soil examinations (HAYES, COLL) 2. The effect of fertilization on the odor and flavor of cigar tobacco (COOLHAAS) 17. Chemical transformations caused by the decomposition of vegetable wastes (VINASSA DE REONY) 8. Rotenone, the effective constituent of derris root (TAKEL, et al) 10. Artificial silk [by product fertilizer] (Ger. pat. 517,747) 23. Preserving bone material (Brit. pat. 340,010) 29.

Agriculture and Live-Stock in India (*New journal*) Published bi monthly by the Imperial Council of Agricultural Research, Govt. of India Central Publication Branch, P. O. Box 2078, Calcutta. Vol 1, No 1 appeared Jan., 1931. Price, Rs 6 or 9s 9d per year.

The Indian Journal of Agricultural Science (*New journal*) Published bi monthly by the Imperial Council of Agricultural Research, Govt. of India Central Publication Branch, P. O. Box 2078, Calcutta. Vol 1, No 1 appeared Feb., 1931. Price, Rs 10 or 16s 6d per year.

Device for estimating the acidity of the soil. FRANZ SCHNURPFEL and ALFONS FREISLEBEN. Ger. 519,368, Oct. 6 1929.

Fertilizers. ARMOUR FERTILIZER WORKS. Brit. 340,120, Dec. 11, 1928. Liquid anhyd. NH_3 is added to fertilizer materials or mixts., at least one of which is acidic, in sufficient quantity substantially to neutralize the free acid of the materials. This treatment e.g., may be applied to a mixt. of superphosphate, KCl or K_2SO_4 , $(\text{NH}_4)_2\text{SO}_4$ and NaNO_3 , with or without addn. of org. materials such as tankage or cottonseed meal. Cf. C. A. 24, 2538.

Fertilizers. ODDA SMELTEVERK AKTIESELSKABET and E. JOHNSON. Brit. 339,969, Nov. 24, 1928. Compd. fertilizers are obtained in processes in which the mother liquor, obtained after the crystn. of $\text{Ca}(\text{NO}_3)_2$ from a soln. of phosphate rock in HNO_3 ,

as described in Brit. 339,310 (C. A. 25, 2514). It is neutralized with or caused to react with NH_3 , urea, cyanamide or cyanamide salts such as Ca cyanamide, inorg. phosphates, compds. of NH_3 and CO_2 , alkali carbonates, etc., so that after removal of water from the product a dry stable fertilizer is obtained. Other fertilizing substances may be added, and various details of procedure, app. used, etc., are given.

Fertilizers. HENRIK WITTEK. Ger. 517,503, Sept. 10, 1925. A non-corrosive non-powdering lime-nitrogen fertilizer is prepd. by forming an aq. soln. of the lime-nitrogen and then adding clear neutral oil while heating, or as vapor. The PhOH - and base-free fraction going over between 180-300 in a stone coal distn. is mentioned as a suitable oil.

Fertilizers. GUSTAV THIER. Ger. 517,527, Jan. 15, 1928. App. for directly azotizing heated CaCO_3 to produce CaNCN is described.

Fertilizer. MASCHINENFAB. A. G. VORMAIS BRÜCK & HENKEL. Ger. 517,551, Feb. 18, 1920. App. for working up animal matter such as flesh and bone is described.

Fertilizer. ERIKA RUDOLF A. HARTMANN. Ger. 517,050, Nov. 3, 1928. Animal, fish etc., waste is extd. with steam and the ext. ppd. with metal salts such as $\text{Al}_2(\text{SO}_4)_3$, $\text{Fe}_2(\text{SO}_4)_3$, etc., or quercitinated or tanning ext. The product is used as a fertilizer.

Fertilizer. PAUL COURTY. Fr. 1,072,270, June 12, 1920. A granular fertilizer contg. NH_4NO_3 , which retains its powd. state during storage is made by adding CaCO_3 or other non-hygroscopic substance to the molten fertilizer, then adding a less amt. in powd. form to the granulated fertilizer.

Fertilizers. F. G. FÄHRNIND A. G. Fr. 1,061,081, May 21, 1920. A fertilizer contg. urea and $\text{Ca}(\text{NO}_3)_2$ is made by melting or fusing together the components almost or completely free from water. A rotating furnace may be used.

Fertilizers. F. G. FÄHRNIND A. G. Fr. 1,071,151, June 7, 1920. A salt contg. $(\text{NH}_4)_2\text{SO}_4$ and NH_4NO_3 in the mol. ratio of about 1 to 3 is crystal. and afterward mixed with addnl. substances to prevent the particles agglutinating, such as kieselsühr, KCl , $(\text{NH}_4)_2\text{HPO}_4$, or $(\text{NH}_4)_2\text{SO}_4$.

Fertilizers. SOC. D'ÉCLAIR CHIM. FOUR C'IE. Fr. 696,874, June 10, 1930. K salts are intimately mixed with phosphates and the mixt. treated with suitable amts. of dil. mineral acids, such as HNO_3 or H_2PO_4 . Cf. C. A. 25, 767.

Fertilizers. SOC. VILAIN FRÈRES. Fr. 1,061,008, April 11, 1930. Concd. Mg fertilizers are made by calcining to redness globulite and like natural carbonates of Mg and then treating with an acid such as HNO_3 , HCl , H_2SO_4 or H_3PO_4 . Cf. C. A. 25, 1225.

Fertilizers. SOC. VILAIN FRÈRES. Fr. 1,067,067, June 3, 1930. A fertilizer is prepd. by the reaction of H_2PO_4 on MgO until complete soln. in an excess of H_2PO_4 is obtained, and then NH_3 is added until the excess of H_2PO_4 is neutralized. The clear liquid is decanted from the ppt. formed and used again. Fr. 1,071,068 describes the prepn. of a fertilizer by adding to the usual reagents for the prepn. of superphosphates MgO and NH_3 , if desired in the form of a double sulfate, or MgO or a salt of Mg alone may be added. Cf. C. A. 24, 4110.

Plant preservation. TISCH & STAMMOW INTERNATIONAL CO. & FR. SCHADLINGS-ABTÄUFUNG M. N. H. Ger. 517,642, Apr. 5, 1930. To avoid the risk of poisoning the plants themselves when using alkylene oxide-contg. gases as insecticides, fungicides, etc., these gases are brought into contact with a heat source to convert the alkylene oxide into aldehyde. Steam may also be formed by the heat source for mixing with the aldehydes.

Insectifuges and insecticides. SCHMIDT KAUFLAUF A. G. (Herbert Schotte and Karl Gornitz, inventors). Ger. 520,330, Nov. 15, 1925. Esters of thiocyanic acid are used, with or without the usual additives.

Insecticides, etc. DEUTSCHE GOLD- UND SILBER-SCHNITTANSTALT VORMALS ROSSLER. Ger. 517,631, July 21, 1925. Insecticides, etc., prepd. consisting mainly of HCN , have a sufficient amt. of chloropicrin present to ensure a warning odor.

Insecticides. KÖHM & HAAS CO. Fr. 1,061,326, May 20, 1930. An insecticide is composed of an aliphatic mono-thiocyanate, the org. group of which contains one or more neg. groups. Examples are given of the use of monothiocyanate of the mono-ethyl ether and the monomethyl ether of diethylenglycol, the first in water and the second in kerosene. Other products are referred to.

Insecticide, etc. ARPAD DE WICZINSZKY. Fr. 1,053,503, May 10, 1930. Hydrogenation products of pyridine and its homologs, particularly piperidine, are used for the destruction of insects. Fr. 1,053,501 describes the use of furalol or substances liberating furalol in the organisms such as furalamide or the disulfite compd. for destroying rodents.

Insecticide. WM. MOORE (to American Cyanamid Co.). U. S. 1,707,877, March

24 An insecticide which is suitable for killing plant lice comprises trimethyl chloroethyl ammonium chloride (suitably used as a spray with various other substances)

Fungicides for seeds. CHEM. FAB. LUDWIG MEYER Ger 521,235, June 28, 1925 The activity of known fungicides is enhanced by addn. of a small quantity of 1 or its compds Cf C A 24, 5928

Fungicides for seeds. I. G. FARBWIND A-G (Max Bockmühl and Walter Persch, inventors) Ger 519,270, Apr 17, 1928 The addn. compds of lig salts with aromatic compds having an unvald side chain are used Suitable compds are obtainable from lig(OAc), or ligCl₂ and allyl phenols or p-nitrophenyl allyl ether Cf C A 25, 1626

Fungicides for seeds. I. G. FARBWIND A-G (Max Bockmühl, Walter Persch, Kaspar Haff and Robert Kramer, inventors) Ger 519,488, Apr 8, 1928 Mercurized thiophene and its homologs are used, e.g., a thiophene mercury acetate

Destroying leaf-cutting ants. MEINHARD JACOBY Ger 516,595, Mar 23, 1929 A prepn. giving off a gaseous or vaporous poison is used The example mentions a mixt. of KCN, NaHSO₄ and water, acrolein or H₂S may be added

Weed killer. CHEMISCHE FABRIK LUDWIG MEYER Fr 695,672, May 14, 1930 A weed killer is composed of a mixt. of CuSO₄ with an alkali chloride or nitrate

16—THE FERMENTATION INDUSTRIES

C. N. FAYE

František Ducháček, JAN SATAVA Chem. Listy 25, 97-101(1931)—A short biography and bibliography of D's contributions to fermentation are given P. M.

Actualities and progress in the distilling industries. FRENCH STEIN Chem. Listy 25, 128-35(1931)—The application of lab. methods to dist. industries, especially to those of Slovakia, is discussed FRANK MARESI

Monograph and analyses of principal alcoholic extracts produced in this country (Mexico). FLORENIO ALVAREZ Rev. quim. 6, No. 8, 7-20(1931)—A states that there are many alc. exts. being manufactured by old and new methods The treatment of maguey by heating and pressing of the stalks and fermentation and distn. of the juices is typical and is given in detail Other sources in addn. to maguey are given The chemistry of the processes to which the juices are subjected is discussed The phys. properties of the finished liquors, the pharmaceutical uses and the chem. analyses are given JOHN M. LABRINO

Continuous fermentation in the production of lactic acid. F. O. WHITTIER AND L. A. ROGERS Ind. Eng. Chem. 23, 532-4(1931)—A method for continuous lactic acid fermentations of the lactose of sweet whey has been developed in the lab. and has been adapted to commercial operation The use of a lactobacillus and a mycoderm gave a yield of 90% of the theoretical based on the lactose originally present

C. N. FAYE

The cause of the low yields of alcohol in the fermentation of Formosan cane molasses. I. N. TAKETOMI AND S. HANAMURA J. Soc. Chem. Ind., Japan 30, 390-6(1927), Mem. Faculty Sci. Eng., Waseda Univ. 1930, No. 7, 129—Low yields of alc. from the fermentation of Formosan molasses are attributed to an insufficient N supply for the growing yeast The yield of alc. was increased from 82-83% to 97-98% of the theoretical yield by adding 0.1 g. of (NH₄)₂SO₄ or (NH₄)₂PO₄ to 100 cc. of the liquid to be fermented II. N. TAKETOMI AND S. ARITAKE Ind. 31, 856-60(1928), Mem. Faculty Sci. Eng., Waseda Univ. 1930, No. 7, 130-1—The low yield of alc. in the fermentation of Formosan molasses is due principally to the presence of glucose, which has reducing properties, and in some cases it is also due to the insufficiency of nitrogenous nutriment for the yeast L. W. T. CUMINGS

Losses of water-soluble phosphoric acid by the clarification of molasses under acid conditions and heat. H. O. HUMMER Brenneri Ztg 47, 169(1930), Cf C A 25, 1623—The influence of the acidity of molasses upon the losses of P₂O₅ is described S. J.

Experiments with sprouting and aegular rye of 1930 crop STAIGER Brenneri Ztg 48, 13(1931)—S. shows that rye which sprouted slightly because of the weather conditions during harvesting could be processed in distilleries without malt S. J.

Use of corn sirup in the spirit industry. C. LUCKOW Brenneri Ztg 48, 16-7(1931) S. J.

Quality of Hungarian hops of 1929 crop I. IVÁNY Kísérlet. Közlemények 33,

238-9(1930) —Averages for 36 samples of hops are: water 7.20, ash 8.67, α -resin 3.42, β -resin 8.84, γ -resin 2.49 and bitter acids 12.26%. S. S. DE FENILY

Resin content and antiseptic value of Hungarian hops. I. IVÁNY. *Külföldi Késkemények* 32, 327-40(1929) —Average values for the 1928 crop are: water 6.55, ash 7.89, α -resin 4.12, β -resin 8.66, γ -resin 2.74 and bitter acids 12.78%. S. S. DE F

Physicochemical investigation on the coloring materials of grapes and red wines. L. CASALE. *Ann. chim. applicata* 20, 559-66(1930) —The various colors in grapes and red wines are related to the pH value of the liquid, the coloring materials (anins) are amphoteric, the isoelec. pt. for the chloride is between pH 5.4 and 5.8. A. W. C.

Apparatus for the determination of fixed acidity in wines. P. BATTÉ. *J. pharm. chim.* [5] 13, 77-81(1931) —Into a conical flask used as a steam generator, a test tube barely touching the surface of the water is fitted through a stopper. Through it another tube (a, connected with b inside the test tube) passes, conducting steam into the test tube through a stopper. The stopper also carries an exit tube bent at a right angle. Tube a has a lateral safety tube, and the end of b barely dips below the surface of the wine (5 cc.) put into the test tube, so as to avoid excessive foaming after the alc. is removed. The volatile acids are expelled after 25 min. heating, and the vol. is about doubled. The fixed acid is then detd. by phthalimosecopy (cf. C. A. 24, 3193). Lactic acid (1.8 g. per l.) proved non volatile since the blue color of water produced by 10 drops bromocresol green (0.04 g. per 100 cc. 11.0 alc., Clark) was not affected by vapors passing through it. With AcOH vapors, the color soon changed to green then yellow. S. WALDBOTT

Determination of carbon dioxide in beer. J. BLOW AND B. KRAUSE. *Wocksch. Brau.* 47, 471-5(1930) —CO₂ is fixed by adding a 30% soln. of NaOH. An aliquot part is then acidified with H₂SO₄ to liberate the gas, which is carried by a current of air (free of CO₂) into 0.1 N Ba(OH)₂. The excess of Ba(OH)₂ is then detd. by titration. The method is described for taking samples from cask and bottle beer. The device used consists of a 10-cc. pipet with a suitable connection to a compressed air container, a tap and a steel point. A soft rubber stopper and clamp to hold the borer in position are auxiliary equipment. S. J.

The protein in beer. I. Kinds of nitrogenous substances in beer. S. AYAL. *J. Soc. Chem. Ind., Japan* 14, Suppl. binding, 72-4(1931) —Expts. were conducted to det. the amt. of total N (crude protein), protein and non protein N in light-colored 2 months old and 24-months old beer. Methods for each detn. are given. It was found that on aging the coagulative N and albumose N increased slightly, the peptone and amino N decreased. The amide N increased slightly. The peptone and polypeptide N comprise 40-43% of total N, albumose 24-27%, coagulative protein 15-17%, amide 9-10% and amino N 5-10%, called as glycecoll. C. N. FREY

Composition of slop char. W. KILP. *Breweri-Ztg.* 47, 295(1930) —The compn. of slop char varies according to the compn. of molasses, the manner of processing and the method of combustion. Three different slop chars were analyzed and the K₂O content was detd. Various results were found: 45-47 (beet molasses), 29.5 (cane molasses, yellow char), 15.5 (cane molasses, black char), expressed in percentages. S. J.

Phosphatide (lecithin) content of compressed yeast. DREHN. *Breweri-Ztg.* 47, 210(1930) —Two yeast samples (74.5 and 69.4% moisture) were extd. with alc. Alc. exts. of 6.26 and 10.17% were obtained with a P content of 2.38 and 2.44, resp., corresponding to 60-62% phosphatides. Phosphatides of 60-62% purity were easily obtained without recrystn. The yeast is emphasized as a source of phosphatides. S. J.

Yeast fermentation. A consideration of the truth of the Arndt-Schulz rule. H. DANNENBERG. *Klin. Wocksch.* 10, 211-21(1931) —Quinine, phenol and HgCl₂ do not accelerate fermentation by yeast. H. EAGLE

Manufacture of alcoholic motor fuel (DIETRICH) 21. Purification of brewery sewage by the activated sludge process (ALI) 14.

Fermentation processes. DEUTSCHE HYDRIERWERKE A-G. Fr. 698,003, May 31, 1930. In biochem. transformations by means of bacteria, e. g., the production of higher alcs. or org. acids, a preliminary fermentation is carried out in worts strongly diluted and with the addn. of an inoculating substance which may be already cultivated on carrier layers. The concn. of the wort is gradually increased by the addn. of worts of high concn. as soon as the fermentation process has reached its greatest intensity.

Fr 698,036 describes a fermentation process in which a bacteria culture is added to the wort at the culminating point of its physiol activity, the wort being poor in nutrients and adverse to the propagation of cells

Alcohol. SOC ANON DES DISTILLERIES DE DIJON SÈCRETS. Fr 696,775, Aug. 1, 1929. Ale of a concn higher than that of the azeotropic ale water mist is obtained directly from fermented wort or dd ale by combination of the ordinary continuous distn with a method using the principles of azeotropism. The heat contained in the vapors from that part of the app where concn of the ale takes place is transmitted wholly and directly by the vapors to that part of the process where the operation is being carried out in the presence of a third substance, and ale of a concn higher than that of the azeotropic ale water mist is obtained at a point where normally abs ale would be obtained if the transmission of heat was by means of a heat exchanger. The ale obtained is afterward completely dehydrated by known means.

Purifying alcohol. FIRMA I. ATRACT. Fr 696,852, Jan. 18, 1930. Crude ale is purified and dehydrated simultaneously by an azeotropic distn, with or without pressure. No purification or elimination of impurities takes place between the dehydration of the crude ale and the rectification of the ale sep'd from the product of the azeotropic distn except a sepn by decantation from the entraining liquid.

Making butyl alcohol and acetone by fermentation. COMMERCIAL SOLVENTS CORP. Ger 520,383, May 14, 1927. See U S 1,668,614 (C A 22, 2235).

Acetone. HOLZVERKOHLEN INDUSTRIE A G. Fr 696,233, May 28, 1930. In the manuf of acetone by the action of steam on LTOH at high trmps in the presence of catalysts such as O compds of heavy metals, the mist of vapors liberated during the distn of worts is used directly or after sepn of impurities of high bp for the ketonization operation. A suitable app is described. Cf C A 23, 3235.

Tartaric compounds. ANDRÉ MONTOUSSE. Fr 697,179, June 10, 1930. Tartaric compds are extd from vinasses, which contain them for the greater part in the form of K bitartrate. The vinasses are boiled in a certain amt of mother liquor, distilled and diluted and the remaining mass is poured into a fresh amt of mother liquor. The temp is raised sufficiently to maintain the bitartrate in soln, while the insol compds are pptd. The liquid is then decanted and crystd. The mother liquor is neutralized for the following operation.

Distilling sulfite liquor spent. ZELLSTOFFFABRIK WALDHOF and MAX GADE. Ger 519,245, Apr. 25, 1929. The first runnings, which contain the aldehyde fraction, are continuously or periodically returned to the vats in which a fresh wort is fermenting or is to be fermented. The spirit obtained from a wort to which the first runnings have been added contains only the normal amt of aldehyde, the aldehyde in the runnings being eliminated in the fermentation.

Dealcoholizing wines, etc. MARC G. COLPEAU and ANDRÉ L. F. CORNILLAC. Fr 696,773, Aug. 16, 1929. Wines and other fermentd fruit juices are dealcoholized by the Charentais process but at a low temp under high vacuum to cause no alteration of the various constituents and by fractional distn to collect separately the head and tail runnings, which are afterwards reincorporated in the liquid to the desired degree. An app is described.

Wine and spirits. FRIEDRICH FAERBER. Ger 516,547, Oct. 21, 1928. App for purifying mash is described.

Metal vat for wine, etc., fermentation processes. LOUIS GRANGES. Ger 516,549, July 19, 1927.

Filter plug for passing purified air and sulfur dioxide into wine casks. K. PRILLINGER. Brit 379,761, Nov. 18, 1929. Filtering layers may comprise sponge impregnated with salicylic acid and $KMnO_4$ coils and wadding also impregnated with the same substances. Various structural details are described.

Beers poor in alcohol. ALBRECHT HASELBACH. Fr 695,591, May 14, 1930. Use is made of a wort with almost double the usual amt of hops. When fermentation is complete starch sirup, 'capillaire' or bonbon sirup may be added.

Hops. GOTTFRIED JAKOB. Ger 516,548, Jan. 23, 1929. Hops are first extd to give a tannin ext which is used in the manuf of mash or as a sprinkling agent. The extd hops are then worked up with clear wort in the usual way. By this method, the max use is made of the hop constituents.

Air heater for drying hops, etc. VIVIAN FLIXINGTON. Ger 520,372, Nov. 9, 1927. Corresponds to Brit. 285,153.

Yeast. AKTIEBOLAGET SEPARATOR. Fr 696,094, May 26, 1930. Wort charged with yeast prep'd in the usual manner is treated in a centrifuge with continuous evacua-

tion of the wort and of the yeast, which is again centrifuged in a bowl of comparatively large diam until relatively dry

Apparatus for washing yeast. I. ANST MÖLLER Ger 521,274, Sept. 27, 1928.

17—PHARMACEUTICAL CHEMISTRY

W O FINEY

The stabilization of medicinal wines. DOMINGOS DE BARROS *Bol assoc brasil pharm* 11, 36 8(1930) Medicinal wines which are perfectly clear when freshly prep'd often become turbid on standing because of the pptn of substances which remained in the supersat'd soln To avoid this pptn it is advised to prep the wines by extg the vegetable drugs with the wine itself as the solvent instead of using a stronger alc medium followed by diln C ABELEDO

Researches on the combustibility of tobacco. IV. The combustibility in relation to the chemical composition of the leaves. R. NACCORINI (WITH G. MINNA) *Boll ist super agraria Pisa* 6, 263 99(1930) In previous papers (*Ibid* 5, 401, 521 and 599 (1929)) an app for the detn of the combustibility of the various types of tobacco was proposed, and this value was studied in relation to the place of the leaves on the tree, and the soil fertilization In the present expts the leaves of Kentucky tobacco (cultivated near Pisa) were analyzed and it was found that the easily combustible leaves contained higher quantities of amide N, citric, oxalic and pectic acids and Fe_2O_3 , while those having lower burning capacity contained higher quantities of nitric, ammoniacal and nicotinic N, sugars, pentosans, fats and waxes, Na O, N_2O , Cl, SiO_2 , SO_2 , and P_2O_5 There is no difference between the relative contents of CaO, K_2O and malic acid

G. A. BRAVO

Preparation of extractum hydrastis fluidum and methods of analysis. Z. CSIKER *Magyar Gyógyszerészeti Társaság Estenője* 5, 73-85(1929) —A report on the known methods of prep'n and on the official detn of the alkaloid content in the drug and the ext.

B. C. A.

Two types of arsphenamine. RUDOLF L. LARSEN *Am. J. Syphilis* 15, 50-7 (1931) —Not only chem, phys and biological tests, but clinical reports indicate that 2 types of arsphenamine are made, a methylated and a nonmethylated type. Clinical reports check the biological tests in assigning more toxicity to the methylated than to the nonmethylated type.

B. C. BRUNSTETTER

Isolation of anonaine from *Anona squamosa* Linn. FELICIANA R. REYES AND ALFREDO C. SANTOS *Philippine J. Sci.* 44, 409-10(1931) —The seeds of *Anona squamosa* Linn., commonly known as "ates," are used to kill parasites of the human body, especially lice The active principle is a base first isolated by Trimourti (*C. A.* 19, 656). In the present paper it is shown to be identical with anonaine obtained from other sources (*C. A.* 25, 705) The powdered seeds (18 kg.) yielded 0.258 g. of the base, m. 122-3°, by the method of ext'n previously described. Hydrochloride m. 237-8° (decompn), chloroplatinate m. 248.5° (decompn)

ALFRED BURGER

The constituents of *Sweetaria chirata*. A. JANAKI RAM *Proc 15th Indian Sci. Congr.* 1928, 160 —*Sweetaria chirata*, a well known valuable Indian medicinal plant, was partially exam'd by Sahar in 1914 and Subramaniam in 1924 By steam distn. of the alc. ext. R. obtained a cryst acid, a phenol and isoamyl alc The residual resin insol in water yielded an acid m. 312° and a neutral substance m. 133° The nonvolatile portion sol in water contained an acid m. 193°

E. J. C.

The constituents of *Plumbago zeylanica* Linn. M. C. TUNNINKATTI *Proc 15th Indian Sci. Congr.* 1928, 163 —*Plumbago*, the active principle of the drug, has been isolated in a pure form It forms orange-yellow silky needles m. 78°. It is sol in most org solvents and in alk solns with formation of a crimson or pink color The combustion and the mol-wt detn indicate an aromatic ring comp'd The deep red color with $FeCl_3$ soln indicates a phenolic OH group in the comp'd The Bz deriv m. 146°. Further phys and chem properties are being studied with a view to det. the constitution.

E. J. C.

Indian eubeb oil. V. P. SHIVKRE *Proc 15th Indian Sci. Congr.* 1928, 168 —The fruits grown in the Mysore Province gave on distn with steam 7.5% of an essential oil The oil contained *d*-sabinene (28%), *d*-terpinene and cineole (15%), *d*-4-terpineol, *d*-terpineol, an unidentified primary alc, *l*-cadimene and another sesquiterpene Another sample of oil was found to contain *dl*-sabinene and the same constituents as the first sample.

E. J. C.

Examination of the active principle of Indian henna. SYED B. ALI *Proc 15th Indian Sci Congr* 1928, 173 —A water ext. is treated with Pb acetate, the Pb is removed by H_2S , and the active principle is oxid. with benzene. It is quite a pure product, for it gave a const. m. p. It is easily oxidizable, the color changes to dark red even on exposure to air. The oxidized product can be easily reduced by Zn and HCl. The formula found differs from those given by previous workers. I. J. C.

Morphine determination in opium. C. G. VAN ARKEL AND P. VAN DER WIELEN *Pharm Weekblad* 68, 309-16(1931) —The Dutch Pharm. method of morphine detn. does not give concordant duplicates when considerable gum is present in the opium sample. A detailed description is given of the previously recommended method (C. A. 24, 5107) which makes use of ale tartaric acid for the extn. Repeated analyses of Persian, Levantine and Macedonian opium by this method showed close agreement. A. W. DOX

Thalleioquin test for quinine. A. WINOGRAD *Chemist Analyst* 20, No. 3, 7(1931) —Dissolve the alkaloid in 50 000 times as much $CHCl_3$. To 0.5 cc. of the soln., add 2 cc. of Br H_2O and 1 cc. of dil. NH_4OH . An intense, emerald green color is given by as little as 0.1 mg. of quinine. W. T. II

Sensitiveness of the thalleioquin reaction. J. LISINBRAND *Arch Pharm* 269, 65-7(1931) —Quinine sulfate is susceptible of detection in a diln. of 1:100 000 (corresponding to quinine 1:140 000) by the thalleioquin test. However, this reaction is highly dependent on the influence of time, furthermore the action of dil. Br- H_2O must be accurately regulated, as also the pH carefully chosen, in order to raise the sensitiveness to the max. The most favorable conditions for highly dil. solns. with 11 ion concn. 10^{-4} to 10^{-3} appear to require 10 to 15 sec. Further expts. are under way. W. O. F.

History of Austrian pharmacopoeias. OTTO ZIEKERT *Pharm Monatshefte* 12, 21, 55-7, 75-6(1931), cf. C. A. 25, 1916 —An address. W. O. F.

Three hundred years with cinchona bark. OTTO ZIEKERT *Pharm Monatshefte* 12, 25-8(1931) —An address. W. O. F.

Examination of galenical preparations in filtered ultra-violet light. L. ZECHNER AND F. GSTRNER *Pharm Monatshefte* 12, 28-31(1931) —The results obtained in the examn. of a series of fluid exts. in ordinary and filtered ultra violet light are recorded. W. O. F.

Occurrence, application and adulteration of Ipecacuanha root. IRANZ BERGFR. *Pharm Monatshefte* 12, 33-4(1931) —A brief summary of the history of this drug in connection with an outline of its microscopical examn. and possible adulterants. W. O. F.

Drug nomenclature of the last forty years. I. HICZOG *Pharm Monatshefte* 12, 60-1(1931), cf. C. A. 24, 3320 —An address. W. O. F.

Characteristica of the valerates and isovalerates. ARNO MÜLLER *Riechstoffind.* 6, 4-7, 30-41(1931) —A compilation of the essential properties of the several members of these 2 groups in connection with their occurrence in natural products. W. O. F.

German Museum and German perfume industry. A. M. BURGER *Riechstoffind.* 6, 25-33(1931) —A brief description of the German Museum of Munich (apparently the largest tech. and scientific museum in the world), including a few typical displays therein relating to perfumes both crude and finished, in connection with data on, and views of, several noted German establishments for the manuf. of perfumes. W. O. F.

Action of bromides on morphine hydrochloride and ethylmorphine hydrochloride. FISCHER AND HUFFMANN *Süddeutsche Apoth.-Ztg.* 71, 131-2(1931) —Solns. of morphine HCl and ethylmorphine HCl yield with bromides the corresponding hydrobromides of these alkaloids, this is similar to the behavior of codeine phosphate. W. O. F.

Pharmaceutical analysis. C. A. ROJANY IX. Pharmacological and biological evaluation of pharmaceutical preparations. JULIUS A. MÜLLER *Pharm Ztg.* 76, 273-6, 286-8(1931) —This paper deals specifically with adrenaline, anthelmintics, eannabis, disinfectants, antienzymes, skin irritants, insect powders, local anesthetics, narcotics, ergot, valerian preps., digitals and laxatives. W. O. F.

Economic drug testing. XII. L. ROSENTHALE *Pharm Ztg.* 76, 288(1931), cf. C. A. 25, 557 —*Estn. of colchicine in colchicum seed* —Shake 0.6 g. of the powdered sample in a small medicine glass with 12 g. $CHCl_3$ and 0.5 g. NH_4OH for 10 min., filter, distil 6 g. of the soln. to dryness, boil the residue with 10 g. H_2O until the vapor no longer blues red litmus paper, then add to the hot liquid 2.5 g. dil. HCl and about 0.5 g. of paraffin, and cool in running water. Filter, wash with water until the filtrate shows no acidity, then dil. to 250 cc. If 10 cc. of this diln. gives with 9 cc. HCl (d 1.19) and

1 cc phosphotungstic acid soln a turbidity after 15 min (changing to a yellow ppt the following day), the sample contains at least about 0.4% colchicine W O F

Estimation of morphine in opium and its preparations. LAJOS DAVID AND JATVAN NOVAK *Pharm Ztg* 76, 289(1931) cf *C A* 25, 557 —The outstanding feature of the procedure advocated by D. and N. consists in the use of MeOH in place of abs 1 tOH in the extn of morphine, on account of its greater soly in the former alc and of less likelihood of extg contaminating substances W O I

Examination of liquor cresoli saponatus D. A.-B. 6, and aapo kalinus D. A.-B. 6. WALTER MEYER *Pharm Ztg* 76, 125 8(1931) cf *C A* 25, 1031 —Recent exams of preps alleged to be official were found to contain the fatty acids of linseed oil instead of the prescribed oil itself The methods followed in arriving at these conclusions are described in detail W O F

Acetyl determinations. HOFFMANN *Pharm Ztg* 76, 320(1931) —Reference is made to a recent paper (cf *C A* 25, 1761) in which the advantages in the use of the Verley and Baling method of acetylation were emphasized, satisfactory values are reported by this method in the examn of 1 tOH, glycerol, iso-PrOH, thymol, salol, β naphthol menthol and santalol The method failed with citronella oil (geraniol) W O E

Evaluation of acetate of alumina solution OTTO SCHMIDTOLA *Pharm Ztg* 76, 112 3(1931) Attention is directed to the pharmacopeial tests for impurities in the official liquor notably Al formate CaSO_4 and K_2SO_4 These tests are discussed W O F

Preparation of homeopathic triturations. A. KERN *Pharm Ztg* 76, 366-7 (1931) —The prepn of, and certain tests on, the strength and stability of homeopathic triturations are discussed in connection with an app (illustrated) to det any catalytic decompn likely to develop in the finished product W O I

Indian kidney tea (Koemis-Koetjing). G. KUELMANN *Pharm Ztg* 76, 367-9 (1931) —The botanical and pharmacol facts of the plant *Orthosiphon stamineus* Blenth are considered in connection with its chem constituents, which at present offer no adequate explanation of the therapeutic (chiefly the diuretic) action of the tea made therefrom W O E

Diacolation of fluidextract of valerian. H. BERGDIEN *Pharm Ztg* 76, 427(1931) —Directions are given for the prepn of this ext, which is regarded as superior to the one commonly made by percolation W O F

"Etrate," a new highly active type of drug. II. Estimation of the extract content of crude drugs. CLERMUS GROSSI *Pharm Zentralhalle* 72, 183-7(1931); cf *C A* 25, 379 —The present study was undertaken to develop a method capable of yielding a max ext corresponding to the true activity of the drug in question After reviewing the methods commonly employed in the prepn of exts, comparative results are reported on 5 g samples under hot aq extn over a period of 6 hrs, which show the highest values thus far obtainable by any extn method The materials examd were cinchona bark, uva ursi leaves, valerian root, rhubarb rhizome and their resp etrates W O E

Histochemical detection of cetic acid in Iceland moss. M. PRONER *Pharm Zentralhalle* 72, 227(1931) —In nearly every case a section immersed 2-3 min in 25% NH_3 soln developed a rose red to red color in the central hyphae, which changed to brown red on treatment with FeCl_3 W O E

Chemical investigations of the tobacco plant. II. Chemical changes that occur during the curing of Connecticut shade-grown tobacco. HUBERT B. VICKERY AND GEORGE W. PUCHER *Conn Agr Expt Sta Bull* 324, 207-40(1931), *Science* 73, 397-9 (1931), cf *C A* 24, 2352-3 —The catabolic changes occurring in tobacco leaves cured for 0, 12, 18 and 51 days were detd by extg the various N fractions, solids, ash, crude fiber, ether sol solids and carbohydrates in ext of the leaves The ext was prepd by boiling the leaves carefully in water contg sufficient H_2SO_4 to maintain a pH of 4.0 Loss of nicotine was thus prevented The residue from the extn was ground and again extd with boiling water A third time the residue was boiled in water, and the exts. were combined and coned *in vacuo* After completion of the curing process approx. 20% of the solids originally present in the leaf disappeared. Over 50% of this loss was in the nitrogenous constituents Chlorophyll destruction was practically complete in 12 days Sol carbohydrates decreased in curing approx 80%; probably most of this loss was CO_2 and water Extensive deamination of the amino acids followed by amide synthesis occurred during curing of the leaf The protein underwent complete digestion to amino acids At least $\frac{1}{2}$ of the potential amino N of the protein that underwent enzymic hydrolysis in the first 12 days of the curing process was further converted to NH_3 and amide N. The approx constancy of the quantities of NH_3 , amide and amino N in the

exts. suggests some sort of equal condition in the relationships between these forms of N. Approx. 75% of the loss of water and of sol. carbohydrate and more than 50% of the loss of org. solids and of the ether sol. constituents occurred during the first 12 days more than 75% of the quantity of protein digested underwent this process in the same period. Far reaching chem. changes set in very shortly after the leaves were detached from the tobacco plant.

C. R. FELLEAS

Utilization of by-products of saccharin manufacture in synthesis of drugs and in medical science. W. HERRZOG. *Chem. Rundschau Mitteleuropa u. Balkan* 7, No. 15, 107-11 (1930) cf. C. A. 25, 42.—The utilization of *p*-toluenesulfonyl chloride and *p*-toluenesulfonamide is described with detailed data from the patent literature.

S. S. DE FINÁLY

Color and ethereal oil content of pulverized cinnamon. I. HORVÁTH AND L. BENEDEK. *Kisérlet. Közlemények* 33, 787-91 (1930).—No regular connection was found between ethereal oil content and ale. ext. content. The ethereal oil content decreased with quantity of waste particles, especially of outer cork layers. Dark powders mostly contained less ethereal oil.

S. S. DE FINÁLY

Synthesis of local anesthetics. FAIGYES KOVÁK. *Matematik és Természettud. Értelen* 46, 348-60 (in German 361-2) (1929).—Some new derivs. of 3-methoxysalicylic acid were produced. Crystals of 2,3 (m H₂N, C₁₁H₂COO)(MeO)C₁₁H₂CO₂Me had anesthetic influence when put on the tongue. The anesthetic influence of 2,3,5 (BrO)(MeO)(H₂N)C₁₁H₂CO₂Me was somewhat weaker.

S. S. DE FINÁLY

Hungarian essential oils. M. JANCSEK. *Mezőgazdasági Kutatások* 2, 153-6 (1929) cf. C. A. 24, 2235.—The following data are given for the corresponding essential oils: *Coriandrum sativum* d₁₅ = 0.8723, $\alpha_D = +10^\circ 35'$, $n_D^{20} = 1.4653$, ester content 72.6%. *Mentha piperita* (1927) d₁₅ 0.9079, $\alpha_D = -25^\circ 5'$, $n_D^{20} 1.4575$, menthol content 69.34%. *Chenopodium ambrosioides var. anthelminticum* (1927) d₁₅ 0.9852, $\alpha_D = -8^\circ 39'$, $n_D^{20} 1.4760$. *Salvia officinalis* (1927) d₁₅ 0.9193, $\alpha_D = +13^\circ 32'$, $n_D^{20} 1.4631$. *Salvia sclarea* (1928) d₁₅ 0.9043, $\alpha_D = -24^\circ 2'$, $n_D^{20} 1.4648$, ester content 70%. *Majorana hortensis* (1928) d₁₅ 0.9058, $\alpha_D = +29^\circ$, $n_D^{20} 1.4791$. *Lavandula vera* (1928) d₁₅ 0.8924, $\alpha_D = -7^\circ 29'$, $n_D^{20} 1.4613$, ester content 62.62%.

S. S. DE FINÁLY

Composition of tobacco smoke and denicotization of tobacco. J. BODNÁR. *Természettud. Közlemények* 62, 397-401, 447-6 (1930).—A review of the present methods for denicotization. Hungarian "Denikó" cigars and cigarettes did not contain less nicotine than ordinary ones. Cotton wool placed in the mouthpiece of cigarettes retained about 50% of the original nicotine content of the smoke.

S. S. DE FINÁLY

The assay of digitalis on frogs. B. BEHRENS. *Arch. exp. Path. Pharmacol.* 140, 237-56 (1929). *Physiol. Abstracts* 15, 400-1.—The lethal dose for slow intravenous infusion of K-strophanthin in frogs (*R. temporaria*) was determined on 149 animals and found to be 0.3264 γ per g. The standard deviation of a single observation was 20.1%. The curve relating dose to mortality would coincide with the curve obtained by Trevan for the subcutaneous injection of digitalis. When K-strophanthin was injected in 10 doses with 24 frogs on each dose, the corresponding curve had a standard deviation for a single observation of 17.1%. When these curves are compared with de Lind's figure of 13.03% for the standard deviation of cats, it is concluded that frogs are more variable than cats. The official method in Germany for determining the lethal dose entails injecting 6 frogs on each of a series of doses. It is difficult to calculate the accuracy of the results to be expected in this case, so an exptl. method was adopted. A representative imaginary population of frogs was constructed from the results described above. The lethal dose for each frog was written on a card, and expts. carried out by drawing the cards from a hat. It was found that with the standard method in which about 7 groups of 6 frogs were injected, 12% of the expts. gave errors of over 10%, and 26% of the expts. gave no definite result at all. A new method of calcg. results is proposed, which diminishes the no. of large errors and enables all the results to be used. By this method when 39 frogs were used the standard deviation of the result was 4.1%. When 26 frogs were used, the standard deviation was 5.8%. If the frogs were divided up so that the dose interval was smaller and there were proportionately fewer frogs on each dose, the change had no definite effect on the accuracy of the test.

G. G.

Assay of *Lobelia inflata*. O. STEFFUN AND V. SVERBY. *Arch. exp. Path. Pharmacol.* 141, 116-22 (1929). *Physiol. Abstracts* 15, 401.—A method is described for the bioassay of *Lobelia inflata*, using a decapitated cat as an exptl. animal. Lobeline causes a rise in blood pressure in decapitated cats. By comparing the effects of lobeline and nicotine and by extirpating the suprarenals it was shown that lobeline raises blood

pressure only by acting on the suprarenals and has no peripheral action. Hence it is suggested that lobeline should be assayed by comparing its power of raising blood pressure with that of a known soln. of adrenaline. G. G.

Chaulmoogra oil from the seeds of *Hydnocarpus heterophylla* and the requirements of the Dutch Pharmacopeia V. D. R. KOOLHAAS. *Pharm Tijdschr Nederland* 1930, 313 7(1930) cf. C. A. 24, 2133. The Dutch and Brit. pharmacopoeias require that chaulmoogra oil be pressed from seeds of *Tarakogenos kurasu*, King. Chaulmoogra oils from several other botanical species are known, with nearly the same chem. const. and therapeutic values. Especially the required soln. in alc. (due to the high acid value) and the high acid value itself (21.27 Brit. Pharm.) are irrational since this oil can be gathered with a low and const. acid value (0.44). The high m. p. (32°) of the fat of *Hydnocarpus heterophylla*, indigenous to Java, the low acid value (<1) and the chem. compn. (glycerides of chaulmoogric and hydnocarpic acids) make this oil a good one for the combating of leprosy in the Dutch East Indies, though it does not meet the requirements of the Dutch Pharm. D. R. KOOLHAAS.

The effect of fertilization on the odor and flavor of cigar tobacco. C. COOLHAAS. *Proefsta. Vorstenland Tabak Mededeel.* No. 66, 1 22(1930). Fertilization with $(\text{NH}_4)_2\text{SO}_4$ plus superphosphate gave tobacco of the best odor and flavor. Tobacco fertilized with $(\text{NH}_4)_2\text{SO}_4$ alone was slightly inferior, while that fertilized with farmyard manure and that grown without fertilizer were much inferior. D. R. KOOLHAAS.

Investigations on fire-holding capacity of tobacco leaf. C. COOLHAAS. *Proefsta. Vorstenland Tabak Mededeel.* No. 68, 1 71(1930). The av. glowing duration was detd. by noting the time in sec. that each leaf burned, when ignited with a glowing carbon point at the bottom of the leaf on the right side between two veins. About 800 leaves per test were burned. The duration of glowing is directly proportional to the logarithm of the percentage of moisture. Previous observations were confirmed that a poorly burning leaf contains more Cl and less K than a leaf that burns well. The contents of Ca, Mg, and citric, oxalic and malic acids also affect the burning quality, but SO_4 , SiO_2 and Na_2O do not. The duration of glowing is not improved by fertilizing; on the contrary, certain fertilizers injure the fire holding capacity. The injurious effect is closely associated with the amount of Cl contained in the org. fertilizers. $(\text{NH}_4)_2\text{SO}_4$, if necessary supplemented with double superphosphate, is detrimental to the fire holding capacity of the tobacco, although less so than stable manure. Potash fertilization did not improve the fire holding capacity, probably because the absorption of the K is limited by the antagonistic influence of Na_2O , CaO and MgO. When a leaf was gathered and the stalk placed in a weak potash soln., a considerable increase of the duration of glowing was recorded, after the leaf had been cured and fermented in the usual manner. Most non burning tobacco was found in places where quantities of Cl existed in the soil moisture. It is therefore probable that all measures to break the soil capillaries will be of value. Different varieties of tobacco grown under identical circumstances showed great differences in the av. duration of glowing. D. R. KOOLHAAS.

Harmine and its preparation from the root *Peganum harmala* L. A. D. ROZENFELD. *Farm. Zhur.* 1930, 181-6. R. reviews briefly the work on harmine, describes its method of prepn. and phys. properties according to Merck and points out the identity of harmine with banisterine. Harmine has always been prepd. from the seed, but R. was successful in prepg. a harmine base from the root of *Peganum harmala* L. in 3% yield. It forms large needle shaped crystals, m. 258°. HCl salt m. 202°. It gives all the reactions common to banisterine and harmine. The water soln. is fluorescent, especially under ultra violet light. B. S. LEVINE.

Potassium ferrocyanide as a microchemical reagent for the identification of strychnine. Yu. D. GRESIN. *Farm. Zhur.* 1930, 291-5. G. describes a drop method on a test plate. Strychnine, codeine and narcotine, when treated with appropriate solns. of HCl and the yellow and red cyanide, yield sp. crystals readily recognizable under the microscope. The crystals are represented in six drawings. B. S. LEVINE.

Botany, chemistry and pharmacodynamics of *Toddalia aculeata*. J. E. LONSTEIN AND P. HESS. *Bull. soc. pharmacol.* 38, 157-64(1931). The root contains a glueoside that is cleavable by emulsin, forms white crystals and is purified by pptn. with $(\text{NH}_4)_2\text{SO}_4$. The oleoresin isolated is yellow and of bitter taste, d₄ 0.873, $\alpha_D^{20} = -15^\circ 30'$, b. 190°. It consists of a mixt. of Me_2O , PhCO_2H , eugenol, citronellol and a paraffin b. 80°. A dark brown resin is contained in a quantity of 1.25%; it is sol. in EtOH , Et_2O , CHCl_3 , C_6H_6 , benzene and turpentine oil, insol. in petroleum ether. Other constituents are: saponifiable lipides 0.37%, in 39.6°, l. value 66.64, n_D^{20} 1.6339, unsaponifiable lipides 0.43% contg. 2.15% phytosterol, reducing sugar, as glucose, 0.46%; hydrolyzable sugar, as sucrose, 0.12%, gum 0.80, protein 5.82, citric, malic, oxalic

and succinic acids. Ca , Mg , Na , I , Mn , Zn , As , Cl , P , S and Si . The resin in doses of 0.3 g. produces abortion in pregnant guinea pigs of 500 g. in non pregnant, contractions of the uterus. Larger doses cause convulsions, paralysis of the posterior part and death within 48 hrs. The leaves contain a gluco alkaloid, *toddalaine*, which is pptd by all reagents for alkaloids, and 2.55% of tannin. In pharmacol expts toddalaine decreases the chronaxia of both nerve and muscle, and disturbs the isochronism. In the isolated heart the speed and the amplitude are reduced. A. I. MEYER

Detection of adulteration of aloes. C. LIGEA. *J. pharm. chim.* [8], 13, 143-4 (1931).—L. recommends a simple soly. test with NH_4OH in the place of the more complex chlorinated aloin test (C. A. 20, 2561). In a conical flask mix 2 g. powd. aloes with 25 cc. of a soln. of 10 cc. official NH_4OH in H_2O to make 100 cc. Cork, allow to stand for 1 hr., shaking 3-4 times, then filter and wash the ppt. with dil. NH_4OH soln. until the filtrate is colorless. Avoid evapn. of NH_3 by covering the funnel to prevent the ppt. becoming gelatinous. Dry it in air at 25-30° and weigh. The results show that Curaçao aloes with 0.25% insol. are superior to all others, confirming the aloin Cl_2 test. A carefully prepd. Ooganda aloes gave 0.60%. A Cape aloes gave 2.40, 9.35 and 3.40%, resp. Cape aloes No. 4 gave 45.5% of an insol. vegetable powder, hence it was adulterated. This method obviously fails to detect resin, if its presence is suspected, ext. with benzene and apply the usual tests. S. WALDBOTT

Biological assay of digitalis by the Hatcher-Magnus method applied to the dog. J. PRIOR. *Thesis*, Paris, 1929, *J. pharm. chim.* [4], 13, 143 (1931).—The substitution of the dog for the cat renders certain changes of procedure necessary. Processes are given to obtain a max. soly. of the active principles of digitalis in infusions and tinctures. Then the min. lethal doses were detd. for several samples of powd. digitalis by means of infusions of various concns. or ale. tinctures. Some expts. consisted in the same detns. after previous injection of a non lethal dose of ouabaine, the constancy of the min. lethal dose of which has been previously studied at different concns. The max. error in all these detns. does not exceed 18%. S. WALDBOTT

The chaulmoogra of Cameroun. R. MATHIAS. *Thesis*, Paris, 1929, *J. pharm. chim.* [8], 13, 183-4 (1931).—The morphological and chem. aspects of *Caloncoba welwitschii* Gilg, the chaulmoogra of Cameroun are described. This shrub and *C. glauca* Gilg show distinct morphological differences in their leaves and fruits. The const. of the oil of *C. welwitschii* are close to those found recently for the oil of *Oncoba echinola* Oliver (cf. C. A. 23, 112, 4297). Figures and micrographic descriptions are given of the grains of various species and of the press cake obtained from each of these. S. WALDBOTT

Experimental investigations concerning the standardization and the pharmacology of heart tonics with a new method. W. NYRT and LOUISA DUBOIS. *J. Am. Pharm. Assoc.* 19, 945-7 (1930). cf. C. A. 25, 1949.—The work is in 2 parts: (a) standardization of heart tonics and (b) pharmacology of digitalis, with special reference to the relationship of Ca and H^+ ions to the drug. The scope of the studies is indicated by the use of 50 cats, 300 rabbits and 4500 frogs. In (a) warm blooded animals are to be preferred. Great variations are found in the frog as a test animal. The no. of animals necessary to make a satisfactory assay of a prepn. of digitalis should be left to the judgment of the individual worker. The formula of van Wijngaarden allows too wide a range of variations and does not fulfil the purpose for which it was recommended. Intravenous administration is best in the assay of heart tonics. The new intravenous anesthesia (barbital Na) is to be preferred over former methods. The drop in blood pressure to 0 approaches closest the theoretical end point and is preferred to the stoppage of the heart. It is very important to keep the time of the individual expt. as uniform as possible. The chief difficulty in standardizing digitalis preps. (for heart use) is that some require preliminary treatment before assay, such as evapn. of the ale., etc. These manipulations frequently ppt. active principles. The basic requirement for a physiol. assay should avoid such damaging treatment. A method has been worked out by N. and D. which meets the principal enumerated requirements. The rabbit is used as the test animal. The advantages of the method are that animal material is always available. The end point is obtained by the drop in blood pressure supplemented by ouabain and is as close as may be expected. Rabbits have a higher resistance to heart tonics than other warmer blooded animals so that tinctures need not be diluted more than 1:4 which does not interfere with the test. Preps. of great concns. as well as those of high dilutions may be tested. (b) Digitalis and strophanthin poisoning of the heart takes place in the absence of Ca^{++} . Increase of Ca^{++} to 4 times the normal reinforces and hastens digitalis action. The acid base equil. of the nutrient fluid of the heart may be changed within the range of pH 5.2 to 7.6. Full digitalis action occurs within the above mentioned range. L. E. WARREN

Studies on strychnine. The relative sensitivity of certain chemical and physiological tests. JUSTUS C WARD AND JAMES C MUNCH *J Am Pharm Assoc* 19, 954-7 (1930) — Analyzed specimens of strychnine salts were subjected to seven of the more common color tests. The most sensitive test is $K_2Cr_2O_7 + H_2SO_4$, by which 1.0 γ can be detected. $KIO_3 + H_2SO_4$ was the least sensitive of the oxidizing tests. Malachon's test was the least sensitive of any tested, 67 γ being required to obtain the reaction. A quant test for strychnine was developed which depends on the sensitivity of the individual to bitter taste in tasting various concns of the alkaloid or its salts. Individuals vary in sensitivity but the av limit is 5 γ . NaCl and sucrose decrease the apparent bitterness. Different areas of the tongue differ in sensitivity. After an individual has been standardized by tasting knowns it is possible to detect differences of 5%.

A comparison of the pH of neoarsphenamine and sulfarsphenamine in relation to differences in their structure. A E JURIST AND W G CHRISTIANSEN *J Am Pharm Assoc* 19, 951-3 (1930) — The pH range of 9 specimens of sulfarsphenamine and 1 specimen of neoarsphenamine acid lay between 2.15 and 4.30. Sulfarsphenamine contains some free acid as was shown by detns of As, S and Na. The excess of S in some specimens is due to the presence of sulfarsphenamine acid. This acid is stable while neoarsphenamine acid is unstable. Conclusion: The observed facts are sufficient to warrant the belief that the structures of neoarsphenamine and sulfarsphenamine are considerably different.

A toxicological investigation of chloral hydrate. L W RISING AND L V LANN *J Am Pharm Assoc* 20, 9-11 (1931) — It is generally believed that $CCl_3CH(OH)H_2O$ (C) is very unstable in the presence of decomg org matter, such as the cadaver, but there are no exptl data. Knowledge is also lacking concerning the effects of preservatives on the stability of the substance. Stomachs of sheep were minced and known weights mixed with known quantities of (C). The preservative was added and the container sealed with paraffin. The preservatives were EtOH, 1% $HgCl_2$, a satd soln of As_2O_3 , 10% HCHO and an embalmer's cavity fluid contg EtOH, MeOH, $C_2H_5(OH)$, H_2O , etheral and other constituents not named. Analyses were made at 1, 3 and 7 months. (C) was detd by steam distn, subsequent reduction with Zn and titration by $AgNO_3$. In most tests there was progressive decompn. At the end of 7 months the undecompd amts were with EtOH 41, $HClO$ 79, $HgCl_2$ 91, As_2O_3 89, cavity fluid 88 and no preservation 63%. R. and L. believe that 18 months would elapse before all of (C) would be decompd. In forensic cases $HgCl_2$ is apparently the best preservative to employ in keeping tissues for analysis for (C), EtOH is the poorest. Analysts should recover at least 50% of (C) if death has occurred within 6 months.

Monograph and analyses of principal alcoholic extracts produced in Mexico (ALVAREZ) 16. The "Deutsche Museum" and its section of pharmacy (SILVA ARAUJO) 2. Effect of C on tobacco culture (BONUCCELLI, CRIV) 15. *Cannabis indica* resin (CAHN) 10. Colored condensation products [intermediates for drugs] (Ger. pat 519,051) 10.

BERNOULLI, EUGEN, and THOMAN, JULIUS. Übersicht der gefährlichen und neueren Arzneimittel für Ärzte, Apotheker und Zahnärzte. 3rd ed., revised. Basel: Schwabe & Co. 577 pp. Linen, M. 8.

PARRY, ERNEST J. Sandalwood Oil. Mysore Gov. Printing Office. 46 pp. Reviewed in *Pharm. J.* 126, 364 (1931).

SLÖTTA, K. H. Grundriss der modernen Arzneistoff-Synthese. Stuttgart: F. Enke. About 200 pp. About M 16, linen, about M 18.

Storing hygroscopic drugs and chemicals. HANS GFELLER. *Swiss* 143,544, Oct 22, 1929. The substances are stored in a container supported in a second vessel contg CaO .

Synthetic drug. I G FARBERIND, A-G (Karl Streitwolf, Alfred Fehrlé, Walter Herrmann and Paul Fritzsche, inventors). Ger. 520,225, Aug 20, 1928. 2,3-Di-methyl-4-aminopyrazolone-1 phenyl-4'-arsonic acid and 3 hydroxybenzaldehyde 4-arsonic acid are brought to reaction, yielding a condensation product which is less poisonous than either initial material. The reaction may be effected by boiling together aq solns of the Na salts of the acids. An example is given.

Synthetic drugs (dinitro- and diamino-pyrazolidones). DEUTSCHE GOLD UND SILBER-SCHNEIDANSTALT vorm ROESSLER (to I G Farberind A G). Brit. 339,932, Sept 29, 1928. Pyrazidone is nitrated to produce a dinitro compd from which, by reduction, the corresponding diamino compd is obtained.

Therapeutic agent PERCY DESUNESS Ger 517,935, April 12, 1928. A stable water sol medicine contg I and Ag is obtained by mixing an aq soln of colloidal Ag with Na 7 iodo 8 hydroxyquinoline 5 sulfonate. The medicine is used against tuberculosis.

Salts of alkaloids with camphoric acid. SCHIRING-KAHLHAUM A-G (Walter Schoeller and Herbert Schotte, inventors) Ger. 519,033, Apr 16, 1929. Acid camphorates of hyoscyamine, scopolamine and atropine are prepd by the customary salt forming methods. 1 examples are given.

Dilithium sodium or potassium citrate. FIRMA JOHANN A WOLFFING Ger 516,672 Nov 19, 1929. $\text{Li}_2\text{NaCaH}_2\text{O}_7$ or $\text{Li}_2\text{KC}_2\text{H}_2\text{O}_7$ is prepd by adding Li_2CO_3 to a soln of citric acid in alc., boiling off the CO_2 , adding NaOH or KOH dissolved in the least possible amt of water to the boiling soln., cooling, crystg and drying. The compds are used in therapy in anti arthritis or anti diabetes prepn. Examples are given. The yield is 95%.

Organic phosphorus compounds. I G FARBENIND A-G (Wilhelm Lommel and Heinrich Munzel, inventors) Ger 521,205, Nov 1, 1928. New compds are prepd by condensing triarylphosphine oxides with phenols. Thus, a condensation product m 149-150° is obtained by stirring tri-*n*-tolylphosphine oxide with an equimol amt of PhOH at 150°. Numerous other examples are given also. The products are useful for combating microorganisms, moths, etc., and as intermediates for drugs. Cf C A 24,4121.

Mercury compounds. CARL L J LÖNNMANN Ger 519,320, Aug 7, 1928. Double compds of HgS with Hg halides or other Hg salts are prepd, in highly dispersed form by the reaction of a colloidal soln of HgS with the other Hg salt. The reaction may be effected by shaking a colloidal soln of HgS with the other Hg salt in the cold, or by boiling the mixt for some hrs. The colloidal soln of HgS may be prepd by treating a soln of a Hg halide with Na₂S and NaI in the presence of a stabilizer. 1 examples are given of the prepn of HgCl_2 , HgS and HgI_2 , HgS. The products are useful for treating syphilis.

Acridine derivatives. I G FARBENIND A-G (Karl Schranz and Hans P Muller, inventors) Ger 520,078 Dec 14, 1928. Double compds of 3,6 diamino-10-alkylacridinium salts with salts of 3,6 diaminoacridine are prepd by bringing the salts together, or by the reaction of one of the bases with an acid salt of the other base. The reaction proceeds in the cold or on warming, and a solvent may be present. Examples are given. The products are more sol in water than their components, and have bactericidal properties.

Derivatives of pyrazolone. CURT RATH and ARTHUR BINZ Ger 516,534, Nov 26, 1924. 1 Pyridyl 3 methyl 5-pyrazolone or its substitution products is treated with an equimol amt or excess of an alkylating agent. Thus, 2 [3'-methylpyrazolonyl 1'] 5-nitropyridine is heated with KOH, I_2 , SO_4 and MeOH to give 2 [2',3'-dimethylpyrazolonyl 1'] 5-nitropyridine, m 172-173°. Another example describes the prepn of 2 [2',3'-dimethylpyrazolonyl 1'] 5-chloropyridine, m 135°. The products are used in therapy.

Quinoline derivatives. JOHANNA WOLFFENSTEIN NER DEEN Ger 520,155, July 8, 1926. 2 Phenylquinoline-4 carboxylic acid and its derivs substituted in the phenyl residue with an amino or a dimethylamino group, are esterified with I_2 , $\text{NC}_2\text{H}_5\text{OH}$ by the customary esterification processes. Examples are given. The products have anesthetic properties resembling those of cocaine.

Salts of quaternary bases. IGON GLÜCKSMANN Ger 519,324, Sept 30, 1927. New salts of halogen substituted aliphatic quaternary bases are prepd by the reaction of known salts of the bases with suitable salts of the acids to be combined therewith. Thus, the halides of the bases are caused to react with Ag salts of the acids, or the sulfates of the bases with Ba salts of the acids. An excess of the acid to be combined may be present. Examples are given of the prepn of β , γ -dibromopropyltrimethylammonium coumarin-3 carboxylate, m about 130°, triethyl β -chloroethylammonium cinnamate, tri methyl β -bromoethylammonium *p*-toluenesulfonate, sodiumethyltrimethylammonium mandelate, β , γ -dibromopropyltrimethylammonium theobromine 1 acetate, m about 150°, sodo methyltrimethylammonium acid phthalate, m about 133°, β , γ -dibromopropyltrimethylammonium acetate, m 120-2°, and trimethyl β -bromoethylammonium propionate, m about 180°. The products are of pharmaceutical value.

Antiseptics. F HOFFMANN-LA ROCHE & Co, A G Ger 519,321, April 8, 1930. Addn to 507,418 (C A 25, 560). Substituted phenoxyphenols or their alkali salts are converted into their carbonates or carbamates by the customary processes. Examples are given of the prepn of hydroquinone *p*-tolyl ether carbamate, m 130-1°, 2-chloro-4-phenoxyphenyl carbamate, m 129°, and thymol *p*-hydroxyphenyl ether carbonate, b₁₁ 350°.

Use of quinoxalines with benzoic acid as bactericidal agents. E SCHULTZE Brit. 339,602, July 22, 1929. Benzoic acid is used together with substances such as eucupino-

toxine or vucimotoxine for sterilizing, disinfecting or preserving various products and materials

Analgeso-hypnotic composition (a white inodorous molecular compound of pyrimidone and ethylisopropylbarbituric acid.) WERNER M LAUTER (to E. R. Squibb & Sons) U. S. 1,797,425, March 24

Thyruis gland extracts. I G JARDFIND A G Brit 339,917, Sept 10, 1929
Thymus gland ext is tested in regard to its physiol activity by detg the smallest dose which, on injection into an immature female animal, will prevent the normal occurrence of rutting. An ext made with Ringer's soln is tested and standardized by diln or by concn in vacuo at low temp, in accord with results of the test

Purifying biocatalysts. KNORR A G CHEM LAB Ger 521,126, Aug 29, 1926
Co-enzymes, larmunes, water sol vitamins, etc, are purified by fractional diffusion or dialysis. Thus, an impure soln of the biocatalyst may be carefully introduced at the bottom of a column of solvent, whereupon the biocatalyst diffuses up the column more quickly than some of the impurities and less quickly than others, so that a relatively pure product is obtainable from an intermediate section of the column. Alternatively, the impure soln may be dialyzed through two membranes, the first of which retains high mol impurities, while the second retains the biocatalyst but is permeable to low mol impurities such as salts. 1 sample is given

Hormones. SIGMUND FRANKEL Ger 519,321, June 28, 1927 See Brit 292,062 (C A 23, 1474-A)

Hormones. DOCA ANON FOUR LIND CHIM A DALE Ger 521,106, Sept 29, 1928 See Brit 298,089 (C A 23, 2786)

Female sexual hormones from vegetable starting materials. SIGFRIED LOWE Ger 517,761, June 16, 1926 Female plant tissues are extd with water or org solvent, and the ext is freed from undesired constituents. The prep produces 'heat' in female mammals. In example, the ovaries of the water rose *Nuphar* and willow catkins are dried and extd with ether. The ext is worked up to give a product which produces "heat" in ovariectomized mice

Removing nicotine from tobacco by prolonged extraction with water. LUDWIG LITPMANN Ger 520,439, Dec. 29, 1926

18—ACIDS, ALKALIES, SALTS AND SUNDRIES

E. M. SIMMONS

Purification of crude sulfuric acid. H A J PIETERS AND M J MANNENS *Chem Weekblad* 28, 130-1 (1931) —Crude H_2SO_4 can be purified by leaching through H_2S or air contg about 1 vol % of H_2S . Arsenic is pptd quantitatively, nitrates are completely and nitrates partly eliminated. The ppt can be removed by filtration or by skim of the acid till the ppt settles. Addn of mineral oil to the acid before pptn causes the ppt. to form at the oil acid boundary and the clear acid can be siphoned off. F DE L

Quinquennial review of the mineral production of India for the years 1924 to 1928. Alum and aluminous sulfates. C S LOX *Records Geol Surv India* 64, 311-8 (1930) —Kalmite, alunite and alunogen occur in India in limited amts. Practically all the alum is produced by the sepn of $\text{Al}_2(\text{SO}_4)_3$ from ironcontg pyritic shales, with the addn of niter or wood ashes. The shale is built up into heaps 18 ft high with layers of wood and clay that had been once burned and exposed to the weather for 1 yr. This is ignited, fresh material added on one side and roasted material removed from the other. The complicated leaching process is described. Alum is being displaced in most of its Indian uses by $\text{Al}_2(\text{SO}_4)_3$ from bauxite. Sulfur, sulfuric acid and soluble sulfates. B. H. PASCON. *Ibid* 440 6 —Small quantities of S occur in several volcanoes. H_2SO_4 is being manufd from Siedian S. Production data are given. The Indian $(\text{NH}_4)_2\text{SO}_4$ industry is considered briefly. ALDEN H EMERY

Preparation of active magnesium carbonate from magnesite. FRANK C. MATHERS AND WALTER V. FACLESON *Proc Indiana Acad Sci* 38, 175-82 (1929) —Pptd MgCO_3 causes lime plaster to set quickly, but mineral MgCO_3 is not effective. The cost of pptd MgCO_3 is prohibitive. The authors prepd active MgCO_3 by calcination of magnesite to MgO , followed by hydration and carbonation. Variables studied were fineness of grinding of original magnesite, temp and time of burning, time of soaking the oxide in water for hydration and time of carbonation. It is recommended that the magnesite be burned at 1000° for 0.5 hr. for pieces passing a 100 mesh screen, and 2 hrs for pieces 4-5 cm in diam, that this oxide be hydrated by soaking for 24 hrs in water, and that

this hydroxide be carbonated by treating with CO_2 under slight pressure with agitation. Pieces of $\text{Mg}(\text{OH})_2$, which are 4-5 cm in diam. require 72 hrs' carbonation; pieces small enough to float will carbonate in 0.5 hr. $\text{Ca}(\text{OH})_2$, contg. 4% of its wt. of this MgCO_3 produces a plaster which sets in 3 hrs. H. M. STARR

Historical notes on salt and salt manufacture. L. G. M. BAAS-BECKING. *Sci. Monthly* May, 1931, 434-46. E. H.

Liquid CO_2 —how technology has harnessed the available sources. GUSTAVE T. REICH. *Chem. Met. Eng.* 38, 130-41 (1931).—The processes for the concn. and purification of CO_2 for liquid CO_2 production are described. Flow sheets of the principal processes are given. L. W. T. CUMMINGS

The water-soluble phosphates—their principal industrial uses. CHARLES H. BUTCHER. *Chem. Trade J.* 83, 217-9 (1931).— Na_2HPO_4 is now replacing, to a considerable extent Na_2CO_3 for softening of boiler and laundry waters, and as a detergent. NH_4 phosphates are used largely in concd. fertilizers. Na phosphates are used as detergents and degreasing agents, and in silk weighting and fireproofing agents nearly all the Na and NH_4 phosphates have been employed. Other applications are mentioned. W. H. BOYNTON

Trisodium phosphate—its manufacture and use. FOSTER D. SNELL. *Ind. Eng. Chem.* 23, 470-4 (1931).— H_3PO_4 is added to Na_2CO_3 soln. to form Na_2HPO_4 . CO_2 escaping. This soln. is filtered, dild. and NaOH soln. added to form Na_3PO_4 soln., which is filtered hot, and crystd. A spray congealed product may be formed if desired. Caking of Na_3PO_4 crystals may be prevented by formation of double salts. The detergent value of Na_3PO_4 is due to its high pH in soln., its ability to lower interfacial tension against oils or solids, and its marked power of emulsifying oils. It is used as a water softener and boiler compd. because of its ability to ppt. Ca and Mg phosphates. E. M. SYMMES

California desert soda—plant of the Natural Soda Products Company at Keeler, California, on Owens Lake. G. ROSS ROBERTSON. *Ind. Eng. Chem.* 23, 478-81 (1931).—Deposits and plant operation are described. E. M. SYMMES

Rate of calcination of limestone. C. C. FURNAS. *Ind. Eng. Chem.* 23, 534-8 (1931).—Calcination of limestone takes place in a very narrow zone, which is the phase boundary between CaCO_3 and CaO . This zone advances from the outside to the inside of the piece at a const. rate for each temp., independently of particle size or degree of calcination. Curves and data are given for rates of calcination and temp. histories of particles. Most of the resistance to heat transfer into the piece appears to be in the narrow zone of calcination, not in the body of the calcined material. The calcination data may be used to det. the surface area of the particles. E. M. SYMMES

Oyster shells as raw material for chemical lime. J. B. NAALEY. *Chem. Met. Eng.* 38, 145-7 (1931).—The only plant using oyster shells as a raw material for producing chem. lime is described. The plant is located at Houston, Texas, and obtains oyster shells by dredging on the Gulf Coast. The principal problem is that of removal of mechanically entrapped foreign material. L. W. T. CUMMINGS

Testing of the decolorizing power of activated carbons. A. S. SIRYAGIN AND E. S. SERKIN. *Zhur. Sakharos Prom.* 4, 466-70 (1930). V. E. BAIKOV

The economics of hydrogen production from ammonia. HANS J. BRAUN. *Metall.-berste* 21, 267-8 (1931).—The Berthier and Burke process for the production of H_2 by the disson. of NH_3 can be operated industrially in Germany at a low cost. Comparative production and shipping costs for both NH_3 and H_2 are given. RUSSELL C. EAB

The present state of bromine technology in the potash mines. J. WALTER AND C. HOEST. *Bull. soc. ind. Mulhouse* 97, 29-44 (1931).—A detailed description is given of the app. and procedure used in the recovery of Br from the potash mother liquors of France and Germany. I. H. ODELL

Method for reclaiming iodine from titration residues. F. TH. VAN VOORST. *Chem. Weekblad* 28, 129-30 (1931).—The I-contg. residues are collected and the I is pptd. as CuI , by addn. of CuSO_4 and NaHSO_4 , the liquid is siphoned off and the process repeated until enough CuI is collected. The CuI is oxidized with crude H_2SO_5 and $\text{Na}_2\text{Cr}_2\text{O}_7$ and the free I distd. off. Liquids contg. CN^- cannot be treated in this way. The difference in the cost of the I and of the chemicals used is large enough to make the method economical. F. DE LEEUW

Hafnium. PAUL M. TYLER. *Bur. Mines, Information Circ.* 6457, 11 pp. (1931).—There are discussed description and properties, occurrence (list of minerals), identification and analysis, geographic distribution, history, sepn. of Zr and Hf , prepn. of Hf and Zr , prepn. of pure Hf salts, and production. ALDEN H. EMBAY

Detection and importance of saponins. L. KOPLER. *Pharm Monatshefte* 12, 53-5(1931)—An address W. O. E.

Laminated material—method of production. HERBERT W. ROWELL. *The Times Trade and Eng Suppl* 28, No 664, 19(1931) E. M. SYMMES

Furfural and its application in the plastic industry. G. GENIV. *Rev gén. mat plastiques* 7, 89-95(1931)—A review of the properties and constitution of furfural and of its com applications as a solvent and as a base for synthetic resins A. P. C.

The importance of the isoelectric point in the manufacture of casein for paper coating. R. W. BELL. *Paper Trade J* 92, No 13, 55-6(1931)—The isoelec. point is at a pH of 4.6, at which point min soly occurs. Casein having this pH value, by whatever process it has been prepd., will, if other factors are properly controlled, be low in ash, dissolve easily, have high adhesive strength and good flowing properties. A. P. C.

Fractional crystallization in the Blanc process with HCl and HNO₃ (MALGRON) 2. Removing PH₃ from gases [with production of gas contg H and N or making NH₃] (Fr pat 696,512) 13

Hydrochloric acid recovery from liquids. GREGORIAN MINKOFF and MARCEL LEVY (to International Sugar and Alcohol Co., Ltd.) U. S. 1,708,099, March 24. See Brit. 307,986 (C. A. 24, 208)

Storing hydrocyanic acid. HEERDT-LINGLER GES. Brit. 339,717, Oct. 22, 1928. HCN (alone or absorbed in porous or granular substances) is packed in closed metal containers from which the atm. O is expelled by boiling the HCN, by generation of an inert gas such as N or CO₂ in the container, or by pumping.

Stabilizing hydrocyanic acid. DEUTSCHES GOLD- UND SILBER-SCHNEIDANSTALT VORMALS ROESSLER. Ger. 517,918, July 4, 1925. HCN is rendered stable by adding such org. halogen compds. (excepting irritants) which split off acid residues under the given conditions. Examples are C₆H₅Cl, C₆H₅Br, C₆H₅I, C₆H₅Cl₂, C₆H₅Br₂, etc.

Nitric acid. NIKOBAUM CARO and ALBERT R. FRANK. Fr. 696,401, Aug. 17, 1929. Gaseous mixts. contg oxides of N, O and steam are obtained by the oxidation of NH₃. A part of the steam is removed, the remaining water is sepd. during a 2nd phase, after or during the oxidation of the gaseous oxides of N, in a 3rd phase N₂O₄ or concd. HNO₃ is obtained. Details of the process are given. Cf. C. A. 25, 2527.

Phosphoric acid. SOC. DES PHOSPHATES TUNISIENS ET DES ENGRAIS ET PRODUITS CHIM. Fr. 696,127, May 26, 1930. Conc'd. H₃PO₄ is clarified by satg. the heated acid with H₂S, adding As₂S₃, stirring and then allowing the solids to ppt.

Phosphoric acid and hydrogen. MAREK LARSSON (to Dul'ont Ammonia Corp.). U. S. 1,797,726, March 24. Phosphide of Fe is treated with steam or steam and air at an elevated temp. to oxidize the P of the phosphide to P₂O₅ and liberate H from the H₂O. An arrangement of app. is described. Cf. C. A. 24, 3090.

Eliminating arsenic from acid to be used in the manufacture of ammonium sulfate. SOC. GÉNÉRALE DES FOURN. À COKE SYSTÈMES LECOCQ. Fr. 696,341, May 31, 1930. See Belg. 361,815 (C. A. 24, 4596).

Sulfuric acid. STANLEY ROBSON and BERTRAM LAMBERT. U. S. 1,797,928, March 24. In the catalytic oxidation of SO₂, sulfurous gases are subjected to such treatment only as is necessary to remove suspended solid matter and then passed in the presence of O over a catalyst contg. solid Cr hydroxide hydrogel contg. at least one adsorbed base.

Sulfuric acid. MAX SCHROEDER. Fr. 695,583, May 14, 1930. Mixts. rich in SO₂ and air are obtained from poor gases such as roasting gases by absorbing the SO₂ in water under pressure and treating this soln. under atm. pressure in a stream action tower with the amt. of air necessary to produce a suitable gaseous mixt. for the production of H₂SO₄. A moderate heating may be used if desired.

Sulfuric acid. NATIONAL PROCESSES, LTD., STANLEY ROBSON and PHILIP S. LEWIS. Fr. 696,476, May 14, 1930. Catalytic materials for the transformation of SO₂ into SO₃ are made by mixing a compd. having a basis of V with an indifferent support in a finely divided state, damping the mixt. and transforming it into an agglomerate with a suitable binder, and treating the damp agglomerate in the cold with a reducing gas until the product has an acid reaction and the V is completely reduced. The V compd. is either vanadic anhydride in the granular state, a metal vanadate or a soln. of a sol. vanadate.

Sulfuric acid. INDUSTRIELEMISKA AKTIEBOLAGET. Fr. 697,173, June 10, 1930. SO₂ is oxidized to SO₃, which is hydrated in the liquid phase by leading, blowing or other.

wise causing the gas to pass through a layer or layers of liquid. The layer of liquid may contain H_2SO_4 and HNO_3 or nitrosylsulfuric acid.

Sulfuric acid. NATIONAL PROCESSES, LTD. Ger 516,764, Sept. 13, 1928. See Brit 303,459 (C A 23, 4539) and Brit 301,833 (C A 23, 4300).

Rotary furnace for concentrating acids. SILESIA, VEREIN CHEMISCHE FABRIKEN (Georg Narten, inventor) Ger 517,935, June 20, 1929. Details of construction are given.

Separating ammonia from gas mixtures. I G FARBENIND A G (Friedrich Frowein and Lrich Rahlfs, inventors) Ger 519,225, Jan 11, 1929. The mixts are treated with mother liquors which contain metal salts capable of combining with NH_3 and which cannot easily be worked up in any other way. A suitable mother liquor is obtained in the manuf. of KNO_3 from $\text{Ca}(\text{NO}_3)_2$ and KCl . The liquor comprises CaCl_2 with some KNO_3 , and when treated while cooling with gases contg NH_3 , e. g., with NH_3 synthesis products, $\text{CaCl}_2 \cdot 6\text{NH}_3$ is pptd. After sepg the ppt., the residual soln may be returned to the first stage of the KNO_3 manuf., or worked up separately to recover KNO_3 . The pptd $\text{CaCl}_2 \cdot 6\text{NH}_3$ is decompd by heat. Cf C A 25, 2250.

Alkali salts. CHEMIEVERFAHREN GES. Fr. 697,069, June 3, 1930. KNO_3 and Na_2CO_3 are prepd. A lye, obtained by mixing CaSO_4 in a mother liquor contg Na_2CO_3 and sepg the CaCO_3 thus formed, is treated with a crude sylvinitic salt of potash and NH_3 , which produces glaserite and KCl . The K_2SO_4 obtained by treating this mixt. with water, is treated with CaCO_3 , previously obtained, and HNO_3 , which gives CaSO_4 and KNO_3 . The lye sepd from the glaserite and KCl is cooled to remove a part of the NH_4Cl and is then submitted to the Solvay process.

Alkaline earth cyanates. I G FARBENIND A-G. Fr. 695,497, May 13, 1930. Alk. earth oxides or carbonates are heated to temps. between 130° and 400° with urea or substances contg it.

Alkaline earth metal chlorides from sulfates. THEODOR LICHTENBERGER and LUDWIG KAISER. U S 1,798,091, March 24. An alk. earth metal sulfate is dissolved in its corresponding melted alk. earth metal chloride, the sulfate is reduced to sulfide (suitably by the action of coke) and the sulfide is converted into chloride by injection of Cl .

Reducing alkaline earth sulfates. WILHELM RASSBACH. Ger 520,458, Feb. 14, 1930. A mixt. of alk. earth sulfate with powd. Al is ignited, e. g., by means of a little powd. Mg . The reaction products are alk. earth sulfide and Al_2O_3 , but some alk. earth oxide and SO_2 may also be formed through the vaporization and air oxidation of some of the alk. earth sulfide.

Process for the manufacture of mono-alkali phosphates. SOCIÉTÉ D'ÉTUDES SCIENTIFIQUES ET D'ENTREPRISES INDUSTRIELLES. Belg 372,395, Sept. 30, 1930. An alkali chloride or a mixt. of alkali chlorides is treated with an excess of H_3PO_4 so as to form MH_2PO_4 , H_2PO_4 , in which M is an alkali metal or NH_4 . This salt is treated with alc., the insol. MH_2PO_4 is sepd., and the H_3PO_4 -alc. mixt. is distd. in presence of H_2O , the recovered alc. being used again in the process.

Carbonates. CHEMIEVERFAHREN G. M. B. H. Fr. 696,396, June 2, 1930. K_2CO_3 and Na_2CO_3 are obtained by mixing a Na_2CO_3 liquor with SrSO_4 or BaSO_4 , sepg the alk. earth carbonate, and treating the lye with crude sylvinitic salt of K and NH_3 , which produces glaserite and KCl . Then the K_2SO_4 obtained by treating this mixt. of salt with water is treated in aq. soln with SrCO_3 or BaCO_3 obtained previously, preferably in the presence of CO_2 , which gives a soln of K_2CO_3 or KHCO_3 and a ppt. of K_2SO_4 and SrSO_4 or BaSO_4 . The K_2SO_4 is sepd. from the SrSO_4 or BaSO_4 by heating with water and the K_2CO_3 or KHCO_3 soln is evapd.

Phosphates. JEAN MARGOLÉS. Fr. 696,138, May 27, 1930. Aq. solns of mono-alkali phosphates are stabilized by adding to the phosphate or to its mixts. with the products resulting from the solubilization of crude phosphates, or to the water used for its soln., an acid, preferably org., which is weaker than the acid used for solubilization of the crude phosphate.

Sulfates. ALPHONSE ZIEREN. Fr. 696,447, May 3, 1930. An app. is described for the uniform distribution of H_2SO_4 in mech. sulfate furnaces.

Thiocyanates. EMIL HEINE. Ger 517,759, Nov. 20, 1928. Thiocyanates of the alkali or alk. earth metals are prepd. by heating S compds. of the metals in the required proportions with CN_2H_2 or derivs. at low temps. between 200° and 500° . S may be added during the reaction if necessary. Thus, BaS , S and CN_2H_2 are heated to about 300° to give $\text{Ba}(\text{CNS})_2$. Also NaSH and CN_2H_2 give NaCNS . Several further examples are given. In all cases the yield is above 90%.

Fluorine compounds. I G FARBENIND A-G. Fr. 695,573, May 14, 1930.

Solns of compds contg F and Al are prepd by the action of solns of $AlCl_3$ on difficultly sol compds of F in aq suspension, in such amts that there are at least 2 equivs of Al for 1 of F. Examples are given of the treatment of fluorspar and Na_2SiF_6 . Complex compds of the formula $AlF_3 \cdot AlCl_3$ are formed.

Compounds containing aluminum and fluorine. I G FARBERND. A.-G. Fr. 695,688, May 15, 1930. Difficultly sol compds are pptd from solns contg $AlF_3 \cdot AlCl_3$ (cf Fr 695,573, preceding abstr) by the addn of carbonates or hydroxides of alkali or alk earth metals to form basic salts of Al or by the addn of alkali fluorides to form double fluorides.

Hypochlorite compounds. THE MATHESON ALKALI WORKS. Fr 695,952, July 4, 1929. See U. S. 1,787,018 and 1,787,080 (C. A. 25, 781).

Nitrogen compounds. SOC DÉTUDES CHIM FOUR L'IND. Fr 695,964, May 22, 1930. Cyanamides and metallic cyanides are treated with gases contg hydrocarbons, N and H at a temp above 200° and below 1000° , whereby the elementary N becomes fixed as N derivative.

Beryllium salts. SIEMENS & HALSKE A.-G. (G Hellmut Fischer, inventor). Ger 520,151, Nov 2, 1928. A soln of an alkali beryllium fluoride or other halide is treated with an alk earth hydroxide or a soln thereof. Practically all the Be is thus pptd as the hydroxide which is then worked up to Be salts.

Rubidium salts. GERIART JANDER. Ger 517,921, Jan 27, 1929. Rb salts are obtained by treating Rb carnallite with silicomolybdic acid, and decomposing the resulting Rb silicomolybdate by heating in a current of HCl , with evolution of $MoO_3 \cdot Cl_2$. The Rb is then leached out from the SiO_2 . An example is given.

Water-soluble hydroxides and carbonates. ALBERT FRITZ MEYERHOFER. Ger 521,430, June 15, 1924. Aq solns of the appropriate fluorides under heat and pressure are caused to react with hydroxides or carbonates of metals forming insol fluorides. The amt of water should not exceed that required to produce a concd soln, and may be less.

Metal carbonyls. I G FARBERND A.-G. (Emil Keunecke, inventor). Ger 520,220, Aug 15, 1928. See Fr 677,548 (C. A. 24, 3329) and Fr 37,284 (C. A. 25, 2528).

Metal carbonyls. I G FARBERND. A.-G. (Leo Schlecht, Walter Schubardt and Emil Keunecke, inventors). Ger 517,831, Mar. 7, 1929. See Fr. 691,100 (C. A. 25, 1042).

Aluminum chloride. I. G FARBERND A.-G. (Johannes Brode and Carl Wurster, inventors). Ger 520,152, Dec 28, 1927. Anhyd $AlCl_3$ is obtained in a stable form suitable for transport by passing $AlCl_3$ vapor into a cooled vessel, suitably into the vessel in which it is to be transported, until a solid block of the salt is obtained. The vessel may be conical and may be lined with a material hindering the adhesion of $AlCl_3$ crystals, e. g., with graphite, soot or talc. Cf. C. A. 25, 1644.

Aluminum chloride. TURE R HAGLUND. Ger 521,339, April 1, 1925. $AlCl_3$ is prepd by the action of anhyd chlorinating agents, at a raised temp, on the Al sulfide melts that are obtained by fusing bauxite or the like with reducing agents and fluxes contg S. The chlorination temp may be sufficient to effect distn of the $AlCl_3$. Suitable chlorinating agents are Cl_2 , HI and chlorides of S and P.

Aluminum fluoride precipitation. JOHN E MORROW (to Aluminum Co. of Am.). U. S. 1,797,994, March 24. Pptn is effected by adding to an AlF_3 soln a small quantity of a finely divided inert material such as alumina, calcined cryolite or finely divided AlF_3 , and subsequently heating the soln.

Alumina from bauxite. WOLF J MÜLLER and HEINRICH HILLER. Fr 695,586, May 14, 1930. See Brit 330,661 (C. A. 24, 6947).

Alumina and phosphorus. ODDA SMELTEVERK A/S and ERLING JOHNSON. Fr. 695,962, May 22, 1930. Al_2O_3 and P or H_3PO_4 are prepd by fusion under reducing conditions of a phosphate ore, such as natural phosphate, and Al compds such as bauxite, Al phosphate, etc., in the presence of compds contg Fe and S, so as to volatilize the P and to produce a slag contg compds of Al and sulfides which is used for the production of pure Al_2O_3 .

Aluminum oxide products from material containing aluminum sulfide. TURE R. HAGLUND. U. S. 1,797,655, March 24. Sulfide-contg material which may be in a solid condition is treated with an oxidizing gas such as air or SO_2 and H_2O to effect replacement of S by O. Cf. C. A. 24, 1473.

Aluminum sulfate. COLLOID-CHEMISCHE FORSCHUNGS A.-G. Swiss 143,696, Jan. 7, 1929. Al contg material such as bauxite, kaolin, etc., is heated with an equal amt. of concd. H_2SO_4 and the $Al_2(SO_4)_3$ so formed exhd. with water. The aq ext. is evapd,

wise causing the gas to pass through a layer or layers of liquid. The layer of liquid may contain H_2SO_4 and HNO_3 or nitrososulfuric acid.

Sulfuric acid. NATIONAL PROCESSING, LRO. Ger 518,764, Sept. 13, 1928. See Brit. 303,459 (C. A. 23, 453⁹) and Brit. 301,533 (C. A. 23, 450⁹).

Rotary furnace for concentrating acids. SILESIA, VEREIN CHEMISCHER FABRIKEN. (Georg Narten inventor). Ger 517,945, June 20, 1929. Details of construction are given.

Separating ammonia from gas mixtures. I. G. FARBEWEND A.-G. (Friedrich I. Brown and Erich Rahls, inventors). Ger 519,225, Jan. 11, 1929. The mixts. are treated with mother liquors which contain metal salts capable of combining with NH_3 and which cannot easily be worked up in any other way. A suitable mother liquor is obtained in the manu. of KNO_3 from $\text{Ca}(\text{NO}_3)_2$ and KCl . The liquor comprises CaCl_2 with some KNO_3 and when treated while cooling with gases contg. NH_3 , e. g., with NH_3 synthesis products, $\text{CaCl}_2 \cdot 6\text{NH}_3$ is pptd. After sepg. the ppt., the residual soln. may be returned to the first stage of the KNO_3 manu., or worked up separately to recover KNO_3 . The pptd. $\text{CaCl}_2 \cdot 6\text{NH}_3$ is decompd. by heat. Cf. C. A. 25, 2230.

Alkali salts. CHEMIEVERFAHREN GEB. Fr 697,009, June 3, 1930. KNO_3 and Na_2CO_3 are prepd. A lye, obtained by mixing CaSO_4 in a mother liquor contg. Na_2CO_3 and sepg. the CaCO_3 thus formed, is treated with a crude synthetic salt of potash and NH_3 which produces glaserite and KCl . The K_2SO_4 obtained by treating this mixt. with water, is treated with CaCO_3 previously obtained and HNO_3 which gives CaSO_4 and KNO_3 . The lye sepd. from the glaserite and KCl is cooled to remove a part of the NH_4Cl and is then submitted to the Solvay process.

Alkaline earth cyanates. I. G. FARBEWEND A.-G. Fr 695,497, May 13, 1930. Alk. earth oxides or carbonates are heated to temps. between 130° and 400° with urea or substances contg. it.

Alkaline earth metal chlorides from sulfates. THEODOR LICHTENBERGER and LEONID KASSER. (S. 1,798,091, March 24. An alk. earth metal sulfate is dissolved in its corresponding melted alk. earth metal chloride, the sulfate is reduced to sulfide (suitably by the action of coke) and the sulfide is converted into chloride by injection of Cl_2 .

Reducing alkaline earth sulfates. WILHELM RASSBACH. Ger 520,458, Feb. 14, 1930. A mixt. of alk. earth sulfate with powd. Al is ignited, e. g., by means of a little powd. Mg . The reaction products are alk. earth sulfide and Al_2O_3 , but some alk. earth oxide and SO_2 may also be formed through the vaporization and air oxidation of some of the alk. earth sulfide.

Process for the manufacture of mono-alkali phosphates. SOCIÉTÉ D'ÉTUDES SCIENTIFIQUES ET D'ENTREPRISES INDUSTRIELLES. Belg. 372,593, Sept. 30, 1930. An alkali chloride or a mixt. of alkali chlorides is treated with an excess of H_3PO_4 , so as to form $\text{MH}_2\text{PO}_4 \cdot \text{H}_3\text{PO}_4$ in which M is an alkali metal or NH_4 . This salt is treated with alc. , the insol. MH_2PO_4 is sepd., and the H_3PO_4 -alc. mixt. is distd. in presence of H_2O , the recovered alc. being used again in the process.

Carbonates. CHEMIEVERFAHREN G. M. B. H. Fr 695,386, June 2, 1930. K_2CO_3 and Na_2CO_3 are obtained by mixing a Na_2CO_3 liquor with SrCO_3 or BaSO_4 , sepg. the alk. earth carbonate, and treating the lye with crude synthetic salt of K and NH_3 which produces glaserite and KCl . Then the K_2SO_4 obtained by treating this mixt. of salt with water is treated in aq. soln. with SrCO_3 or BaCO_3 obtained previously, preferably in the presence of CO_2 which gives a soln. of K_2CO_3 or KHCO_3 and a ppt. of K_2SO_4 and SrSO_4 or BaSO_4 . The K_2SO_4 is sepd. from the SrSO_4 or BaSO_4 by heating with water and the K_2CO_3 or KHCO_3 soln. is evaporated.

Phosphates. JEAN MARGOLIS. Fr 696,138, May 27, 1930. Aq. solns. of mono-alkali phosphates are stabilized by adding to the phosphate or to its mixts. with the products resulting from the solubilization of crude phosphates, or to the water used for its soln. an acid, preferably org., which is weaker than the acid used for solubilization of the crude phosphate.

Sulfates. ALPHONSE ZIEFFER. Fr 696,447, May 3, 1930. An app. is described for the uniform distribution of H_2SO_4 in mech. sulfate furnaces.

Thiocyanates. EDM. HENRI. Ger 517,769, Nov. 20, 1928. Thiocyanates of the alkali or alk. earth metals are prepd. by heating S compds. of the metals in the required proportions with CN_2H_2 or derivs. at low temps. between 200° and 500° . S may be added during the reaction if necessary. Thus, BaS , S and CN_2H_2 are heated to about 300° to give $\text{Ba}(\text{CNS})_2$. Also NaSH and CN_2H_2 give NaCNS . Several further examples are given. In all cases the yield is above 90%.

Fluorine compounds. I. G. FARBEWEND A.-G. Fr 695,572, May 14, 1930.

Solns of compds contg F and Al are prepd by the action of solns of $AlCl_3$ on difficultly sol compds of F in aq suspension, in such amts that there are at least 2 equivs of Al for 1 of F. Examples are given of the treatment of fluor spar and Na_2SiF_6 . Complex compds of the formula $AlF_3 \cdot AlCl_3$ are formed.

Compounds containing aluminum and fluorine. I. G. FARBENIND. A.-G., Fr. 695,688, May 15, 1930. Difficultly sol compds are pptd from solns contg AlF_3 , $AlCl_3$ (cf. Fr. 695,573, preceding abstr.) by the addn. of carbonates or hydroxides of alkali or alk. earth metals to form basic salts of Al or by the addn. of alkali fluorides to form double fluorides.

Hypochlorite compounds. THP. MATHIPSON ALKALI WORKS. Fr. 695,952, July 4, 1929. See I. S. 1,787,049 and 1,787,080 (C. A. 25, 781).

Nitrogen compounds. SOC. DÉTUDES CHIM. FOUR L'IND. Fr. 695,961, May 22, 1930. Cyanamides and metallic cyanides are treated with gases contg hydrocarbons, N and H at a temp. above 200° and below 1000°, whereby the elementary N becomes fixed as N derivs.

Beryllium salts. SIMONS & HALSEK A. G. (G. Hellmut Fischer, inventor). Ger. 620,151, Nov. 2, 1928. A soln. of an alkali beryllium fluoride or other halide is treated with an alk. earth hydroxide or a soln. thereof. Practically all the Be is thus pptd. as the hydroxide, which is then worked up to Be salts.

Rubidium salts. GIEHART JANDORF. Ger. 617,921, Jan. 27, 1929. Rb salts are obtained by treating Rb carnallite with silicomolybdic acid, and decomposing the resulting Rb silicomolybdate by heating in a current of HCl , with evolution of MoO_3Cl_2 . The Rb is then leached out from the SiO_2 . An example is given.

Water-soluble hydroxides and carbonates. ALBERT PAITZ MEYERHOFF. Ger. 621,470, June 15, 1921. Aq. solns. of the appropriate fluorides under heat and pressure are caused to react with hydroxides or carbonates of metals forming insol. fluorides. The amt. of water should not exceed that required to produce a concd. soln., and may be less.

Metal carbonyls. I. G. FARBENIND. A.-G. (Emil Keunecke, inventor). Ger. 620,220, Aug. 15, 1928. See Fr. 677,519 (C. A. 24, 3329) and Fr. 37,281 (C. A. 25, 2528).

Metal carbonyls. I. G. FARBENIND. A.-G. (Leo Schlecht, Walter Schubardt and Emil Keunecke, inventors). Ger. 617,831, Mar. 7, 1929. See Fr. 691,100 (C. A. 25, 1012).

Aluminum chloride. I. G. FARBENIND. A.-G. (Johannes Brode and Carl Würster, inventors). Ger. 620,152, Dec. 28, 1927. Anhyd. $AlCl_3$ is obtained in a stable form suitable for transport by passing $AlCl_3$ vapor into a cooled vessel, suitably into the vessel in which it is to be transported, until a solid block of the salt is obtained. The vessel may be conical and may be lined with a material hindering the adhesion of $AlCl_3$ crystals, e. g., with graphite, soot or talc. Cf. C. A. 25, 1614.

Aluminum chloride. TURF R. HAGLUND. Ger. 621,379, April 1, 1925. $AlCl_3$ is prepd by the action of anhyd. chlorinating agents, at a raised temp., on the Al sulfide melts that are obtained by fusing bauxite or the like with reducing agents and fluxes contg S. The chlorination temp. may be sufficient to effect distn. of the $AlCl_3$. Suitable chlorinating agents are Cl_2 , HCl and chlorides of S and P.

Aluminum fluoride precipitation. JOHN I. MORROW (to Aluminum Co. of Am.). U. S. 1,797,994, March 24. Pptn. is effected by adding to an AlF_3 soln. a small quantity of a finely divided inert material such as alumina, calcined cryolite or finely divided Al , and subsequently heating the soln.

Alumina from bauxite. WOOLF J. MÜLLER and HEINRICH HILLER. Fr. 695,580, May 14, 1930. See Brit. 330,661 (C. A. 24, 5947).

Alumina and phosphorus. OODA SMELTEVERK A/S and ERLING JOHNSON. Fr. 695,942, May 22, 1930. Al_2O_3 and P or H_3PO_4 are prepd by fusion under reducing conditions of a phosphate ore, such as natural phosphate, and Al compds. such as bauxite, Al phosphate, etc., in the presence of compds. contg Fe and S, so as to volatilize the P and to produce a slag contg compds. of Al and sulfides which is useful for the production of pure Al_2O_3 .

Aluminum oxide products from material containing aluminum sulfide. TURF R. HAGLUND. U. S. 1,797,555, March 21. Sulfide contg. material which may be in a solid condition is treated with an oxidizing gas such as air or SO_2 and H_2O to effect replacement of S by O. Cf. C. A. 24, 1473.

Aluminum sulfate. COLLOID CHEMISCH. FORSCHUNGS. A.-G. Swiss 143,696, Jan. 7, 1929. Al contg. material such as bauxite, kaolin, etc., is heated with an equal amt. of concd. H_2SO_4 and the $Al_2(SO_4)_3$ so formed extd. with water. The aq. ext. is evapd.,

purified by addn. of alc., filtered and the $Al_2(SO_4)_3$ obtained by crystn. An example is given stating that a 50-60% yield is obtained.

Ammonium sulfate. DANIEL VORLÄNDER and ALBERT LAINAU. Ger. 519,048, Aug. 22, 1929. $(NH_4)_2SO_4$ is prepd. by oxidizing $(NH_4)_2SO_3$, or a mixt. of NH_3 and SO_2 , in the presence of a complex cobalt sulfite-ammonia compd., which may be formed *in situ*. Thus air may be passed over solid $(NH_4)_2SO_3$ moistened with an aq. soln. of a cobaltamine or an ammoniacal soln. of $CoSO_4$ may be circulated in countercurrent to a mixt. of air, NH_3 and SO_2 . Gases poor in SO_2 may be treated, and the air used may be completely deoxygenated in the process.

Ammonium sulfate. GEWERKSCHAFT VICTOR STICKSTOFFWERKE (Hans Schmalfeldt inventor). Ger. 517,967, May 7, 1929. In evapn. $(NH_4)_2SO_4$ solns., Na_2SO_4 and NH_4HCO_3 are deposited as well as $(NH_4)_2SO_4$. To avoid this, the soln. is neutralized with HNO_3 prior to evapn. so that $(NH_4)_2SO_4$ can be crystd. out at a certain stage of the evapn. leaving NH_4NO_3 in soln.

Ammonium carbonate. KALI CHEMIE A.-G. (Friedrich Rusberg and Gustav Claus inventors). Ger. 517,758, Mar. 2, 1926. Addn. to 488,757 (C. A. 24, 2251). The crust of solid $(NH_4)_2CO_3$ formed on the walls of the chamber, previously moistened with water or $(NH_4)_2CO_3$ soln. into which gaseous NH_3 and CO_2 are led, as described in 488,757 is improved by keeping the temp. of the chamber below 60° .

Antimony oxides. DEUTSCHE SCHMELZ- UND RAFFINIERWERKE A.-G. Fr. 696,448, May 5, 1930. Pure amorphous oxides of Sb, particularly the trioxide, with good covering power are made by heating Sb, Sb ores or mixts. of ores to a high temp. and then burning the metal with air or O_2 dild. with inert gases, e. g., O_2 and CO_2 .

Barium carbide and acetylene. INTERNATIONAL INDUSTRIAL & CHEMICAL CO., LTD. Ger. 517,993, Nov. 29, 1928. See Fr. 682,479 (C. A. 24, 4362).

Calcium hypochlorite. THE MATHIESON ALKALI WORKS. Ger. 517,964, May 24, 1927. $Ca(OCl)_2$ is formed by chlorinating an aq. lime mash and pptd. with $NaCl$. Examples are given. Cf. C. A. 25, 781.

Dehydrating calcium phosphate. I. G. FARBENIND. A.-G. Fr. 697,287, June 12, 1930. $CaHPO_4$ contg. water of crystn. is dehydrated by suspending it in an aq. liquid, preferably of acid reaction, and heating to the temp. necessary for dehydration. Substances favoring dehydration such as fluoindes or silicofluoindes may be added to the liquid. The liquid is used again.

Storing liquid carbon dioxide. P. STAFF. Brit. 333,777, Dec. 3, 1929. A tank and compressor system is described in which liquid CO_2 may be stored under low pressure and near triple-point temp. and various details of operation are given.

Carbon disulfide. I. G. FARBENIND. A.-G. (Friedrich Legler and Hermann Koh, inventors). Ger. 521,336, Dec. 14, 1926. An app. for superheating the S vapor is located in the same heating space as the reaction retort. In the case of an electrically heated retort, the superheater may be located in the retort. Cf. C. A. 25, 2251.

Carbon disulfide. I. G. FARBENIND. A.-G. Ger. 521,337, June 8, 1923. Superheated S vapor is led to a retort contg. heated C. The supply of S vapor may be regulated so that the reaction temp. is maintained in the retort by the heat of the reaction.

Chlorine dioxide. WILLY BECHER. Ger. 517,755, Sept. 27, 1929. In the prepn. of ClO_2 from chlorate and concd. H_2SO_4 , the chlorate is mixed with water and indifferent inorg. material to form solid pieces, into which the concd. H_2SO_4 is then poured. Thus $KClO_3$ is mixed with gypsum made into a paste with water, and dried. 96% H_2SO_4 is then added. Further examples are given.

Magnesium chloride. I. G. FARBENIND. A.-G. Fr. 697,285, June 12, 1930. Hydrated $MgCl_2$ is dehydrated, without much decompn., into MgO and HCl , by introducing it into a mixt. of chlorides in soln., e. g., KCl 50, $NaCl$ 40 and $MgCl_2$ 10%; the content of $MgCl_2$ in soln. not being allowed to pass that of natural carnallite. The dehydration may be combined with the production of Mg by electrolysis by introducing the hydrated $MgCl_2$ into a chamber sepd. from that of the electrodes, in a part of the electrolyte poor in $MgCl_2$.

Magnesia. RHEINISCH WESTFÄLISCHE KALKWERKE. Ger. 519,420, Sept. 27, 1928. See Brit. 319,690 (C. A. 24, 2555).

Potassium nitrate. WINTERSHALL A.-G., CARL T. THORSELL and AUGUST KRISTENSSON. Ger. 517,919, Mar. 17, 1927. Solns. of KCl and HNO_3 are mixed in the concn. described in Ger. 242,014. After sepg. the KNO_3 from the mother liquor, the $NOCl$ present is reduced to NO and worked up to HNO_3 . $FeCl_3$ is used as the reducing agent at low temps. An example is given.

Sodium phosphate. I. G. FARBENIND. A.-G. Fr. 696,360, May 31, 1930. Hydrates of Na_2PO_4 having various amts. of water of crystn. are made by neutralizing

H_2PO_4 with Na_2CO_3 and NaOH in a single operation. The concn of the H_2PO_4 is detd by the content of water of crystn to be obtained. It is treated with an amt of Na_2CO_3 calcd to produce Na_2HPO_4 and an amt of NaOH necessary to form Na_2PO_4 .

Sodium sulfide. I G FARRINGTON A-G. Fr. 696,451, May 6, 1930. Pure anhyd Na_2S is prepd by reducing Na_2SO_4 in molded pieces below the m p of Na_2SO_4 by means of H_2 or other reducing gases. Cl. C. A. 25, 1936.

Sulfur dioxide. SOC ANON POUR L'IND CHIM A HAIE. Brit. 339,926, May 15, 1929, Fr. 695,500, May 13, 1930. SO_2 is obtained in coned form from gases contg it by absorbing it in a mixt of water and an aromatic amine such as aniline, and expelling the absorbed gas by heat.

Furnace for making sulfur dioxide by burning hydrogen sulfide. C. OTTO & Co. G. M. B. H. Ger. 520,150, April 4, 1930.

Aqueous dispersion of titanium dioxide. CHARLES DE ROHDEN (to Commercial Pigments Corp.) U. S. 1,797,760, March 21. For producing an aq soln contg highly dispersed TiO_2 , about 1 part of TiO_2 which has been calcined after coagulation from an aq colloidal dispersion is ground with about 1 part of a dil NH_3 soln and water is added to the resultant paste.

Borax. HENRY BLUMENBERG, JR. Fr. 695,274, May 7, 1930. Borax contained in borax ores is dissolved out with water, the soln is sep'd and a sol comp'd of an alk earth metal is added to ppt an insol borate which is calcined.

Active carbon. I G FARRINGTON A-G (Alwin Mittasch and Josef Jannek, inventors). Ger. 517,966, Feb. 6, 1923. Lignite consisting mostly or wholly of humus sol in aq alkali or NH_3 soln, is heated to 1000° or over. The porous active C so obtained is washed with HCl and is useful in the catalytic prepn of H_2S . Cl. C. A. 25, 2253.

Active carbon; fuel gas. METALLGES A-G (Gustav Bailleul, inventor). Ger. 520,381, Oct. 5, 1929. Cheap adsorption C and a heating gas are obtained by treating carbonaceous materials such as semi-coke and lignite with a mixt of steam and O_2 with or without other gases, contg not more than 6% by vol of O_2 . The steam-gas mixt is preheated to a temp between 120° and 200° , and the reaction temp, which is generally maintained without external heating, is about $500\text{--}600^\circ$. The conditions are adjusted so that the loss in wt of the carbonaceous materials does not exceed 40%. The process may be effected at atm, raised, or reduced pressure, and the carbonaceous materials may be stationary or moved in countercurrent to the steam gas mixt. Examples and addnl details are given.

Compressing chlorine. KREBS & Co., Ges. Brit. 339,855, Oct. 10, 1929. Cl_2 to be compressed is cooled by liquid Cl_2 introduced and allowed to expand in it. App. is described.

Hydrogen. ALFRED T. LARSON (to DuPont Ammonia Corp.) U. S. 1,797,426, March 24. Steam and CO_2 in a volume ratio of at least 1 to 1, are subjected to the action of a catalyst contg Cu and ZnO at a temp sufficiently high (suitably about $300\text{--}600^\circ$) to effect reaction with production of H_2 . Cl. C. A. 25, 1315.

Hydrogen. D. TYLER and IMPERIAL CHEMICAL INDUSTRIES, LTD. Brit. 340,050, Oct. 3, 1929. H_2 or gas mixts contg H_2 are obtained by the alternate passage of hydrocarbon gas (with or without addn of steam) and air through a bed of highly heated coke or refractory material, the air being supplied at different levels within the bed or in different airts at different levels. Various details of app and procedure are described. Cl. C. A. 24, 5117.

Phosphorus. METALLGES A-G. Fr. 696,510, May 27, 1930. P or O compds of P are prep'd from crude phosphates in pieces in a shaft furnace, a phosphate previously litted in pieces in an appropriate app, e. g., a Dwight Lloyd machine is used.

Phosphorus. WSEKOUSNOR OBYEDINENIK KHIMITSCHESKOY PROMISCHLENNOSTI "WSECHIMPROM". Ger. 520,382, Apr. 25, 1930. Pulverulent red P is prep'd by simultaneously heating and grinding yellow P , e. g., in a ball mill. App. is described.

Purifying sulfur. I G FARRINGTON A-G. Fr. 696,010, May 23, 1930. S is purified by treating it at ordinary or raised temp. with acids of medium concn while eliminating, if necessary, the tarry impurities by a simultaneous treating with energetic oxidizing agents.

Crystalline substances. SUPERFINE CHEMICALS, LTD. Fr. 695,508, May 13, 1930. Cryst substances, particularly salts, are obtained in a more easily sol form than the ordinary cryst form by producing, in a vessel having a smooth and substantially continuous interior surface, a relatively large amt of a supersatd soln of the substance in question by cooling in the state of rest a hot coned soln, and at the same time pre-

venting evapn at the surface of the soln. The cooled soln. is then agitated or seeded. Examples are given.

Sulfonation. PAUL BALMGARTEN. Ger 519,046, Aug 12, 1926. Org or inorg substances to be sulfonated are treated in aq. soln. or suspensions with a trisubstituted sulfamic acid prepd by treating a tertiary amine with SO_3 or a compd giving rise thereto. Examples are given.

Active masses suitable for use as catalysts. FAITZ STOEWYNER (to I. G. Farbenind. A. G.). U. S. 1,797,894, March 24. After prepg an irreversible colloid of the second kind such as silica sol it is mixed before drying with a catalytic substance such as CuSO_4 and the mixt is then dried and most of the catalytic substance may be removed (as by acids) to leave a highly adsorbent product. Cf C. A. 24, 3615.

Urea-formaldehyde condensation products. KUNSHARZFABRIK P. POLLAK G. M. B. H. Brit 379,601, May 11, 1929. See Fr 680,110 (C. A. 24, 3867).

Urea-formaldehyde condensation products. FABRIQUES DE PRODUITS DE CHIMIE ORGANIQUE DE L'ATRI J. MALLET and R. ARMENACET. Brit 340,114, Dec 6, 1928. Condensation of urea or its derivs with formaldehyde or its polymers is effected in the presence of a Zn or Pb carbonate or oxide, Bi carbonate, a colored metal carbonate, or TiO_2 (an example being given of the use of ZnCO_3 in producing a moldable product from urea and formaldehyde).

Condensation products containing halogen. I. G. FARBENIND. A. G. Fr 695,602, May 14, 1930. Aq. solns of CH_2O satd with hydrazide are caused to react on phenols with the exception of nitrophenols, halogen phenols and those contg a COOH or aldehyde radical. Cryst. or resinous products are obtained according to the conditions of working and contain halogenated methyl groups. They are useful as starting materials for the manu. of dyes, artificial resins, tanning materials, etc.

Moldable phenolic pentosan material. ORLAND R. SWENNEY (to Iowa State College of Agriculture and Mechanic Arts). U. S. 1,797,559, March 24. Reaction is effected between a solid pentosan-contg material such as corneobs and a phenol (the latter being in a proportion of approx. 60% by wt. of the pentosan-contg material) in the presence of HCl approx. 10% by wt. of the pentosan material, at about 110° for 3 hrs.

Moldable phenolic pentosan material. ORLAND R. SWENNEY (to Iowa State College of Agriculture and Mechanic Arts). U. S. 1,797,593, March 24. A process similar to that described in U. S. 1,797,559 (preceding abstr.) is employed for effecting reaction of ground corneobs and cresol in the presence of S monochloride.

Aromatic condensation products. COMPAGNIE NATIONALE DE MATIÈRES COLORANTES ET MANUFACTURES DE PRODUITS CHIMIQUES DU NORD RÉUNIS ÉTABLISSEMENTS KUEHMANN. Fr 695,980, May 22, 1930. Products which as salts or "free" are emulsifying, wetting or dispersing agents are prepd by condensing aromatic hydrocarbons, sulfonated or not, with esters of CHHSO_3 . Condensing agents such as H_2SO_4 , oleum or CHHSO_3 may be used. Thus, $\text{C}_{12}\text{H}_{11}$ is sulfonated and the reaction product of SO_3Cl on iso- PrOH is added and the mixt. heated to about 100° for 4 hrs.

Celluloid substitute. MOTOMU SUGATA. U. S. 1,797,806, March 24. Camphor (suitably in the proportion of about 50% the wt. of the resin) is added to resin and the mixt. is heated to fuse it to a sticky condition and is mixed with a product such as may be obtained by boiling floss-silk, other silk or like fiber in dil. alkali soln. so as to swell and gelatinize it, suspending homogeneously in alc. or ether and mixing intimately with MgCO_3 or MgO , and heating and kneading.

Artificial substances. BAKELITE CORP. Fr 695,599, May 14, 1930. Nonfibrous artificial substances are composed of viscose, the surface of which carries a coating of a flexible resinoid of phenol and possesses a dielec. resistance of more than 1000 v. per mm.

Artificial substances. HANS KAPPELER. Fr 697,169, June 10, 1930. Infusible substances are prepd by condensing furfural or its derivs with other aldehydes or their polymers with or without solvents or diluents and in the presence of acid condensing agents. Several examples are given.

Artificial masses. I. G. FARBENIND. A. G. Fr 696,008, May 23, 1930. Artificial masses are made by condensing polyvinyl esters with unsatd aldehydes in the presence of catalysts. The products may be vulcanized in the usual manner or fillers, dyes, etc., may be added.

Plastic masses from ivory-nut material. EMIL TELTSCHIK & Co. Ger 516,589, Dec 11, 1927. The ivory nut material is ground up, treated with an org. solvent and added to binding material. Thus, the powd. nut is mixed with S and MgO and treated with CaH_2 . Rubber soln. may be added and the whole pressed at 70 - 100° . Another example describes the admixt. of acetone, celluloid waste and rubber soln.

Moldable compositions. **BAKELITE CORP.** Fr 695,699, May 15, 1930. A moldable compn contains granular or fibrous filling substances and resinoid binders in colloidal dispersion in a soln so as to be easily sepd from the liquid and applied to the filler. Examples are given.

Molded fibrous materials. **COMPAGNIE GÉNÉRALE D'ÉLECTRICITÉ.** Fr 696,962, May 13, 1930. In making molded fibrous materials, the binding agent and the other ingredients of the mixt except the fibers are reduced to a powd state and are dispersed along with the fibers in water. The aq suspension is sepd into an aq and a solid phase, the latter serving, after drying, as a primary material for molding.

Construction material. **FRIDA ROSE REIMANN NÉE VAN JACKSTEN.** Fr. 696,509, May 20, 1930. Coumarone resin is melted with filling materials with the addn of lithopone and if necessary hard waxes, hydrocarbons of high mol wt or rubber.

Artificial horn. **INTERNATIONALE GALALITH GES. HOFF & Co.** Ger 521,246, Jan 14, 1929. Artificial horn is prepd by moistening casein until its total moisture content is about 20-25%, allowing it to stand in the cold for 6-15 hrs., homogenizing it between heated rolls under high pressure and then cooling one of the rolls so as to form a skin. The product is cut off, molded and further worked up as usual.

Adsorption and filter masses. **RHEINHOLD & Co., VEREINIGTE KIESELGUHR- UND KORKSTEIN GES.** Ger 517,756, Feb 23, 1929. Kieselguhr contg Fe and S compds is heated with natural or org addns. The temp is kept below 600° to avoid the formation of readily sol S and Fe compds. Oil or vegetable fiber is mentioned as the added material.

Improving the filtration and decolorizing properties of kieselguhr. **RHEINHOLD & Co., VEREINIGTE KIESELGUHR- UND KORKSTEIN GES.** Ger 517,757, May 7, 1929. The kieselguhr is mixed with natural org material and subjected to a heating process at suitable temps for suitable times. Tables showing the effect of heating various kinds of kieselguhr at temps ranging from 300° to 600° with methylene blue are given.

Kieselguhr. **RHEINHOLD & Co., VEREINIGTE KIESELGUHR- UND KORKSTEINGESELLSCHAFT.** Ger 517,536, Apr 27, 1929. The acid content of kieselguhr is reduced by heating the crude material with exclusion of air, at suitable temps, *e.g.*, 250-600°.

Kaolin purification. **E. MÖLLER.** Brit. 340,142, Jan 4, 1930. Kaolin (previously sepd from raw material, as by fractional pptn with or without the addn of electrolytes, centrifugal action, etc.) is formed into a suspension of a sp gr of 1.06-1.12 and is subjected to further fractional pptn. by allowing it to stand for about 12 hrs. (the height of the suspension vessel being not greater than $\frac{3.2}{S-1}$ cm, where *S* represents the sp gr. of the suspension. The poured-off and filter-pressed suspension is a highly plastic kaolin of a high degree of purity).

Treating bauxite. **METALLGES. A-G (Hans Siegens, inventor).** Ger 509,131, June 2, 1926. In the manus of Al₂O₃ and ferrosilicon by electrothermic reduction with C of the oxide impurities in bauxite, the sepd ferrosilicon is freed from Al by using it to pre reduce a further batch of bauxite. After sepg the purified ferrosilicon, the pre-reduced bauxite is further reduced with C. Operative details and examples are given.

Asbestos and other artificial fibers. **WERNER LÜDKE.** Fr. 697,254, June 12, 1930. Asbestos and like artificial fibers are made by spinning, according to the processes used in the artificial silk industry, sol silicates or mixts of silicates in soln through appropriate coagulating agents such as inert gases or through vapors of org liquids, or org liquid, such as acetone or alc. Neutral salts such as MgCl₂ or CaCl₂ may be used as pptn or conversion agent.

Lime kilns. **LÉANDRE RIGAUD.** Fr 695,343, May 9, 1930.

Apparatus for slaking lime. **SPOERRI & Co.** Swiss 142,389, Dec. 23, 1929. Details are given.

Rotary-drum apparatus for slaking lime. **SPOERRI & Co.** Ger. 519,007, Dec. 28, 1929.

Adhesive. **MARIE A MAJOR NÉE ODAU.** Ger. 521,199, Mar. 17, 1928. A thick adhesive for patching fabrics, etc., comprises crude crepe rubber 10, C₂HCl₄ 80, C₂Cl₄ 20, AcOEt 25 and methylated spirit 15 parts. Cl. C A 25, 1348.

Adhesive. **NAAMLOOZE VENNOOTSCHAP W. A. SCHOLTEN'S CHEMISCHE FABRIEKEN.** Fr 696,635, June 4, 1930. A dry stable product capable of use as a strong glue is made by incorporating sol salts of naphthemic acid with an alk. starch capable of swelling in cold water or like starchy materials.

Viscous adhesive. **ADOLF MÜNGER (to I. G. FARBENIND. A.-G.).** U. S. 1,798,097, March 24. Nitrocellulose and ethyl acetanilide are used together with a softening

agent such as castor oil and a solvent and ground asbestos (suitably at least 10%) to form a viscous highly sticky product suitable for uniting leather or wood veneer to iron or Al. Various other mixts. are described.

Cold-water paste. HENKEL & CO. G. M. & H. Fr. 606,867, June 10, 1930. Starch and products resembling starch are made suitable for the manifold of cold water pastes by heating them under pressure.

Marine algae. COMPAGNIE FRANÇAISE DE L'IODRE ET DE L'ALGUNE. Fr. 604,778, Sept. 6, 1929. Marine algae are submitted to the action of solns. of alk. earth salts which transform the algine into an alk. earth compd. not capable of swelling in water. After soln. with water the compd. is treated with dil. H_2SO_4 , washed and taken up by an alk. soln. to form a soln. of an alkali alginate which is filtered to sep. cellulose. Cf. C. A. 24, 2563.

Cleaning metal surfaces. C. F. DIXLEY. Brit. 340,047, Sept. 28, 1929. Rust and oil, etc., are removed from metal surfaces such as those of vehicle bodies by the use of rust solvents (such as phosphoric, selenic, oxalic, acetic, tartaric or citric acids) together with a compatible oil and grease-solvent such as alcoh., ketones such as Et Me ketone or esters, and a thickening agent (such as raw sienna or metallic oxides such as ferric or Mn oxides) which is inert at ordinary temp. but reacts with the rust solvent on heating or drying and thus renders the compn. friable and "self peeling." A pigment such as willow charcoal or lampblack may also be added.

Polish for lacquered surfaces, etc. ARTHUR VAN DER BERGGEN. Ger. 519,084, May 2, 1928. Collodion, wax Ph_3NO , and petroleum are used with or without acetone.

Composition for cleaning and polishing painted or other surfaces. JAMES G. FANCI. U. S. 1,708,073, March 24. Phosphate rock powder 16, soap 4 and alc. 8 parts.

Coating compositions. JEAN L. SICARD and ALICE SICARD née PARIS. Fr. 687,120, June 6, 1930. A coating compn. which agglomerates instantly is made by heating 3 parts of glue with 1 part of alum to 60° and incorporating 10 parts of saw dust or the like.

Coating materials. HERMANN MAYR. Fr. 687,211, June 11, 1930. A metallic coating on objects of all kinds is obtained by applying several layers of a mixt. of metallic powder, powd. graphite, prepd. chalk, Java chalk, Na_2SiO_3 , and casein. Cf. C. A. 25, 484.

Decorating agents. JEAN DE GRANVILLE. Fr. 606,776, Sept. 4, 1929. A double oxysulfate of Zn and Al is prepd. by the reaction of $\text{Al}(\text{SO}_4)_3$ with ZnO under pressure and at 50° . A plastic mass, which is resistant and very hard, is obtained by carrying out the reaction in the presence of a mass capable of independent crystn. such as reheated CaSO_4 . The mass obtained may be molded, and nitrated cellulose previously dissolved in acetone, etc., may be added during mixing. Metal powders may also be added.

Peroxide containers. I. G. FARBEREND A-G. Fr. 605,408, May 13, 1930. H_2O_2 or other peroxide compds. are produced, treated or stored in metal vessels coated with an enamel very resistant to acids.

Wetting agents. RÖHM & HAAS CO. Fr. 606,328, May 30, 1930. For wetting, emulsifying, or penetrating agents use is made of salts of tertiary alkyl amines in which the alkyl radicals consist of a long straight chain of C atoms and 2 straight chains with not more than 4 C atoms in each, except the hydrochloride of cetvdiethylamine, the remaining valencies of the C atoms being completely sat'd with H. Examples are given of the prepn. of salts of *n*-octyldiethylamine, *n*-butyldiethylamine, dimethylecetylamine and diethyloctadecylamine.

Wetting, etc., agents. I. G. FARBEREND A-G. Fr. 603,814, April 12, 1930. Sulfonic acid esters of aliphatic, cycloaliphatic or aliphatic-aromatic compds., each atom of which has not more than 2 valencies bound to the O, or their salts are used as wetting, cleansing, dispersing or solvent agents. Thus cetyl alc. in ether soln. is treated with ClHSO_3 and the ether is removed, giving a product for use as wetting, etc., agent in the textile industry. Several other examples are given. Fr. 603,815 describes the prepn. of wetting, etc., agents by transforming alcoh. or amines or their derivs. with carboxylic acids or their derivs. into esters or amides, the substances participating in the reaction (aliphatic, cycloaliphatic, aliphatic-aromatic or aromatic) are chosen so that the product of reaction contains at least one aromatic radical, sulfonic groups are introduced. Thus stearic acid and $\text{C}_6\text{H}_5\text{CH}_2\text{OH}$ are boiled while passing a current of HCl gas through. Excess $\text{C}_6\text{H}_5\text{CH}_2\text{OH}$ is removed and the product sulfonated. Several other examples are given.

Wetting and emulsifying agents. H. TH. BÖHMER CHEM. FABRIK. Fr. 606,104, May 26, 1930. Wetting and emulsifying agents are prepd. by condensing triglycerides,

particularly castor oil, or unsatd fatty acids, or oils contg. a large proportion of acids and emulsifiable with water, in the presence of condensing agents, with hydrocarbons (aliphatic or aromatic, satd or unsatd) or with mixts contg hydrocarbons such as solar oil, paraffin oil or resin oil or with appropriate derivs of hydrocarbons, *e. g.*, (1) halogen compds of hydrocarbons such as CHCl_3 , (2) alcs or ethers, (3) aliphatic carboxylic acids, (4) primary or secondary amines or other N compds, except those which form unsol compds with the triglycerides or their fatty acids, and (5) esters and ester acids of alcs. The products may be sulfonated to increase the wetting, etc., properties. Several examples are given. *Cl. C A* 25, 2531.

Transfer sheet. HELENA S. SADLER U S 1,797,997, March 21. A design is printed in sol dyes on a base such as paper and after the design is dried the design-bearing surface is coated with mucilaginous material such as gum tragacanth or gum karaya.

Impregnated fibrocement plates, tubes or other objects. G. ARNOUITS and M. BLAMPAIN Belg 373,033, Sept 30, 1930. The objects, after oven-drying, are impregnated with a mixt of pitch, solvent naphtha and wood oil, or of bitumen, benzene and CCl_4 .

Sound record compositions. A. D. LITTLE, INC., and H. J. BILLINGS. Brit. 340,099, Nov 15, 1929. Outer playing surfaces of records are formed of fibers and a binder of thermoplastic material such as shellac, with a laminated core formed of fibers and a binder including hardwood pitch and preferably also including a resin having low susceptibility to temp changes such as copal resin, manila gum, ternati gum or the like. Numerous details and formulas for prepg the compns used are given.

Bearings. FREDERICK H. RAGAN (to Columbia Axle Co.) U S 1,797,444, March 24. Bearings are formed with a rigid plate or tube having cemented to it a substantial layer of granular compressible material such as cork and graphite.

Bearing material. BEAT H. MCQUEER (to National Carbon Co.). U S 1,797,833, March 24. Bearings are formed comprising carbonaceous material such as petroleum coke, a B compd. such as borax and a polyhydric alc. such as glycol.

Brake linings, etc. I. G. FARBERND A-G. Fr 606,705, June 6, 1930. Fibrous materials are soaked with the products resulting from the condensation of urea, thiourea or their derivs or mixts with aldehydes or their polymers and molding under heat and pressure, if necessary in the presence of small quantities of an acid condensing agent. The condensation may be carried out on the fiber itself.

Dental cements. HANS HEYDEMANN Ger 520,138, Sept. 30, 1927. In the manuf. of these by prepg, oxides, phosphates or silicates of one or more metals of the second or third periodic group by a sintering or fusion process, and then mixing the product with Li_2PO_4 , a compd of Li is added to the material to be fused or sintered. A lower temp. can then be applied, without increasing the soly of the final product.

Cigaret tips. CHEMISCHE FABRIK VON HEYDEN A-G. Ger. 516,752, Dec 31, 1926. Cigaret tips consist of a layer of cellulose hydrate mixed with mineral colors, cork or wood dust, mica, etc.

Pest-destroying agents. HEERDT LINGLER G. M. B. H. Ger. 516,594, Feb. 14, 1928. The agent depends for its action on the evolution of HCN from alk. earth cyanides when exposed to air. The alk. earth cyanides are powd., optionally mixed with irritants and odorants, and pressed into briquets, etc., with binding material. Thus, $\text{Ca}(\text{CN})_2$ is mixed with nicotine sulfate and briquetted.

Phenylurethan derivatives. F. HOFFMANN LA ROCHE & Co A-G. Ger. 519,226, Nov. 23, 1928. The dialkylaminophenyl esters of alkylcarbamic acids are converted into quaternary salts by treatment with dialkyl sulfates. Examples are given. The products are non hygroscopic, stable in air, sol in water, and strongly poisonous. They are useful for destroying animal pests. *Cl. C A* 21, 962.

Rat poison. I. G. FARBERND A-G (Karl Streitwolf, Alfred Fehrle and Hubert Oesterlin, inventors) Ger. 517,513, Sept. 9, 1928. A poison for rodents consists of org aromatic As or Sb compds mixed with fats, waxes, starch, albumin, etc. Bait may also be added. In an example, fresh margarine is melted with chlorophenylarsine oxide. Several further examples are given.

Foam-producing compositions for extinguishing fires. MINIMAX A-G. and HANS BURMEISTER Ger. 531,271, July 31, 1925. A hygroscopic substance, *e. g.*, anhyd Na_2CO_3 , is included in foam producing compns. that are to be dry-stored.

19—GLASS, CLAY PRODUCTS, REFRACTORIES AND ENAMELED METALS

G E BARTON, C H KERR

The history of Hungarian glass manufacture. O KNAPP *Termesztud Közlöny* 61, 46-8(1929) L SZATIMÁRY *Ibid* 491-6—Historical review S S DE FINÁLY

Manufacture and characteristics of laminated glass WILLARD L MORGAN, *Ind Eng Chem* 23, 503-4(1931)—The early work of Benedictus in 1911 is described. In modern manu glass sheets are cut to size, thoroughly cleaned and sprayed with a microscopic layer of gelatin. Several sheets of celluloid, previously sprayed with a mixt of a plasticizer and high boiling solvent, are placed on top of the glass followed by a final glass cover. These sandwiches, so-called, are pressed at 200-240 °F at 150-350 lb pressure for 6-20 min. The edges of the finished glass are sealed with a water proofing substance such as pitch. The use of cellulose acetate instead of celluloid as the inner component of laminated glass has been unsatisfactory. The major problems in the industry are dirt elimination, securing of permanent adhesion, selection of a plastic and control of the laminating process to prevent off color, brittleness or bubbles. JOHN M LADINO

Permanent marking of glass and porcelain vessels. H G ISBELT *Chemist-Analyst* 20, No 3, 20(1931)—A good ceramic ink can be prepared from 1 g K_2CO_3 , 1 g borax 2 g PbO and 2 g $Co(NO_3)_2$. Mix with raw linseed oil with possibly a little turpentine apply with a pen and heat. W T II

Removing the iron of Hungarian sands. L L DE PUTNOKY AND B ROSEF *Magyar Munka Ept Közlönye* 65, 63-70(1931)—Five Hungarian sands were examd for Fe content and means for reducing it to an amount which does not prevent the use of the sand in glass, ceramic, etc., industries. By washing sands with 2 N HCl about 25% of the original Fe was removed. HCl (0.5 N) removed only 10-12% of Fe. HCl also dissolves $CaCO_3$, thus sands contg lime lose much wt and contain relatively even more Fe after washing than originally. HCl attacks the limonite crust of sand particles, but it cannot dissolve grains of siliceous Fe. On washing with water, a sample from Mountain Mátra gave fairly pure quartz, contg 0.2% Fe O_2 and 93.6% SiO_2 with some Al_2O_3 and no CO_2 . Other methods for Fe removal are under examn. S S DE FINÁLY

The effect of the H ion concentration in the adsorption of ions of Ba, Al and Th by clays (HIELD, SOKOLOVA) 2. Plating glass and clay (WERNER) 4. Discoloration of glass by β and γ rays (HOFFMANN) 3. Ceramic plates for lining dyeing and bleaching vats, etc. (Ger pat 520,375) 25.

Apparatus for feeding molten glass. SOC. ANON MANUF DES GLACES ET PRODUITS CHIM. DE SAINT GORAIN, CHAUNY ET CIREY Fr 697,054, June 2, 1930

Glass-melting tank furnaces. W T BARKER (to Hartford Empire Co) Brit. 340,110, May 20, 1929 Structural features

Glass-tank furnace. OTTO ROSENSTEIN (to Illinois Pacific Glass Corp.). U. S. 1,797,500, March 24 Structural features

Manufacture of glass tubes and rods. N-V MAATSCHAPPIJ TOT EXPLOITATIE VAN UIRVINDINGEN Brit 340,676 Oct. 21, 1929 Mech. features

Glass sheets or plates. C LOCKERIE Brit 339,910, March 23, 1929 Various details of app and operation are described for making sheets with an impressed pattern on one side and a fire polished surface on the other

Take-off leering conveyor for sheet-glass-forming apparatus. EUGENE GENTIL (to American Bichroux Co) U. S. 1,797,649, March 24 Structural features

Flating furnaces for reheating glass sheets or plates, etc. FELIX J MALHERBE U. S. 1,797,780, March 24 The life of fireproof materials such as those of the soles of flating furnaces is increased by impregnating them with a mixt of Na and K silicates in petroleum

Apparatus for conveying intermittently rolled plate glass to a number of leers. SOC. ANON DES MANUFACTURES DES GLACES ET PRODUITS CHIMIQUES DE ST GORAIN, CHAUNY & CIREY Ger 521,299, Feb 5 1929

Roll-conveyor leer for intermittently rolled plate glass. N V MAATSCHAPPIJ TOT BEHEER EN EXPLOITATIE VAN OCTROOIEN Ger 521,298, May 11, 1930

Plant for cutting plate glass in plastic state. N-V MAATSCHAPPIJ TOT BEHEER EN EXPLOITATIE VAN OCTROOIEN Ger 519,477, Oct 4, 1928

Glass-annealing apparatus. CHARRÉ & HEUZE (to American Bicheron Co.). U. S. 1,707,650, March 21. Structural features.

Tempering glass. SOC ANON DES MANUFACTURES DES GLACES ET PRODUITS CHIMIQUES DE ST GOBAIN CHAUNY ET CREIL. Brit. 310,108, March 13, 1929. See Fr. 680,890 (C. A. 25, 790).

Apparatus for tempering glass. SOC ANON DES MANUF. DES GLACES ET PRODUITS CHIMIQUES DE ST GOBAIN CHAUNY ET CREIL. Fr. 690,819, Sept. 21, 1929.

Glass permeable to ultra-violet rays. F. WOELNER, R. WOELNER and F. NIXON. Brit. 314,001, March 12, 1929. See Austrian 120,671 and Fr. 691,805 (C. A. 25, 2259).

Treating silver mirrors. RUDOLF RITSCH. Ger. 520,023, June 22, 1929 and 520,021 Oct. 13, 1929 within to 520,021. The mirror layers are mechanically and optically improved, and rendered more resistant to chem. influences, by treating them with HNO_3 vapor (520,021). The same results are attained by treating the mirrors with other oxidizing agents e. g., O_2 , O_3 or H_2O_2 during or after their prepn., the formation of an oxide layer (520,021) being avoided.

Apparatus for rolling and wire-reinforcing glass sheets. CHANCE BROS. & CO., LTD. and A. L. LUKSTER. Brit. 339,561, Nov. 20, 1929. Structural features.

Compound glass. GEORGE I. HAY and MORRIS GRISWOLD. Ger. 519,029, Aug. 30, 1928. A sheet of cellulose acetate or other cellulose deriv. is interposed, while immersed in a softening liquid such as tricresol, between two glass sheets, and the resulting compound sheet is pressed, rolled and finally united by steam pressure. The edges of the sheet may be treated with rubber soln. before the steam treatment.

Composite glass bars. EDWARD J. CONLEY (to Kipling Glass Co.). U. S. 1,797,855, March 24. A glass rod is placed within a glass tube and the intervening space is filled with a liquid cement.

Composite sheets of glass with intervening cellulose derivative material. CLAUDE CHAP. Brit. 339,762, Nov. 21, 1928. Intervening sheets are used for uniting glass, formed of a cellulose deriv. such as cellulose nitrate together with a non-volatile plasticizer such as dibutyl tartrate, tricresol or the like, and stabilizers such as urea, and dyes (preferably of violet line) also may be added. Various other details of method of the composite sheets are also described.

Uniting sheets of glass and cellulose acetate, etc. SOC DES USINES CHIMIQUES REIMS-POUILLEY. Brit. 339,832, April 13, 1929. The surfaces to be united are coated with an adhesive (such as gelatin and celluloid) and the sheets are assembled at slight pressure in a bath of volatile liquid which is a solvent for the plastic sheet only when hot (such as a soln. of ethyl acetate or alc. and water with cellulose acetate sheets) and the materials are then heated in an adhesive coating gas under pressure. Cf. C. A. 24, 1031.

Safety glass. ROBERT P. F. V. HAMON. Fr. 690,767, Sept. 10, 1929. Safety glass is made with sheets of glass united by a layer of transparent or translucent synthetic resins, particularly by products having a basis of urea CH_2O .

Safety glass. I. G. FARMING A. G. Fr. 695,370, May 8, 1930. A layer or layers of esters of polyvinyl alc. such as polyvinyl acetate is interposed between 2 or more layers of glass.

Bleaching and purifying clays and other minerals. R. PRUD. Brit. 339,797, Dec. 18, 1928. "Laths" clay, china, clay, quartz, kalspar, alumin., baryta, Ca and Sr carbonates and sulfates, etc., are bleached and purified by initial washing with peptizing of the colloidal matter by a nonelectrolytic org. substance preventing coagulation such as an infusion of plants, straw or bark, followed by solubilizing of impurities such as Fe by a suitable oxidizing, reducing or oxidizing and base treatment (as by successive use of Cl and SO_2 and NaHSO_3 in the presence of glycerol, glucose or Na hydro-sulfate) and the purified material may be neutralized li. achi. with water contg. lime, Na zincate or a silicate. Various details and modifications of procedure are described.

Ceramic materials. M. LARIVY. Brit. 339,770, Jan. 10, 1929. See Fr. 680,207 (C. A. 25, 672).

Apparatus for making ceramic products in a continuous manner. WILLY LINGENSDORFF and MAX LINGENSDORFF. Fr. 695,059, May 22, 1930.

Apparatus for drying ceramic products. MAISON F. FOUCHE. Fr. 690,618, June 1, 1930.

Drying chamber for porcelain goods. GERNERD NETZSCH. Ger. 521,099, Dec. 22, 1928.

Furnace for baking tiles, etc. SOC SENDORFF-HARFRED GEORGE. Fr. 690,210, May 29, 1930.

- Automatic stoker for brick kilns. ROBERT K. GOODWIN. U S 1,797,608, March 24
- Structural features
- Glazed silica articles. QUARTZ & SILICE SOC. ANON. Ger 519,363, June 3, 1927
- See Brit 273,626 (C A 22, 2040)
- Molded articles of fused silica. QUARTZ ET SILICE Brit. 340,167, April 18, 1929.
- A silica blank is placed over a mold (which may be formed of graphite or Ni) so that an air tight joint is formed between the mold and blank, and suction is produced between the blank and mold while heat is applied to the blank (suitably by an O-H blow-pipe)
- Articles made from mica sheets. I G FARRENTD A-G Fr 696,196, May 23, 1930
- Articles such as insulators are made by uniting sheets of mica by the condensation products of dicarboxylic acids of ethers or thioethers with polyhydric alcs with or without other substances. Examples are given of the use of the condensation products of diglycolic acid or thoglycolic acid and glycerol
- Electrical insulation formed by vitrification of aluminum oxide. SIEMENS & HALSKE A G Brit. 339,726, Oct. 17, 1928
- To avoid formation of Al carbide in producing insulation such as that for spark plugs, by vitrification of Al_2O_3 the heating is carried out in an elec. furnace free from C parts and in a vacuum or atm. free from reducing gases or vapors contg C. An atm. of mixed N and H may be used. Cf C A 24, 447.
- Refractory furnace linings. T H GRAY. Brit. 340,141, Jan 1, 1930. A furnace lining is made of fireclay mixed with about 1.5 times as much foundry sand, which is screened and dampened and made into a paste with the clay. Other fillers may be added.
- Apparatus for projecting refractory material on to furnace walls. BLAW-KNOX Co Ger 519,163, Feb. 23, 1929
- Refractory and acid-proof masses, coatings, etc. STETTNER CRAMOTTE-FABRIK A. G VORM DIDIER. Ger 521,350, Dec. 23, 1924
- Use is made of refractory and acid-proof substances, e. g., Zr or ZrO_2 , with the use as binder of colloidal ThO_2 in the form of a mineral glue obtained by boiling Th salts with acids
- Abrasives. SOC. ANON. DES MANUFACTURES DES GLACES ET PRODUITS CHIMIQUES DE St GOBAIN, CHAUNY & CREY Ger 521,124, June 17, 1928
- Ground bauxite is mixed with up to 2% of an alkali or alk. earth fluoride or fluosilicate, and the mixt. is heated to 1300-1350°
- Enamels. DEUTSCHE GOLD- UND SILBER-SCHNEIDANSTALT VORM ROESSLER. Fr 695 609, May 14, 1930
- Enamels, which are resistant to acids, are made by melting together Li_2SiO_3 or a eutectic mixt. of Li_2O and SiO_2 with a flux such as minium.
- Rotary furnace for enameling and glazing. MITTELDEUTSCHE STAHLWERKE A-G (Wilhelm Philipp inventor) Ger 496,973, Feb. 23, 1928
- Structural details are given.

20—CEMENT AND OTHER BUILDING MATERIALS

J. C. WITT

- Cement investigations in Japan. C. R. PLATMANN *Tonind-Ztg* 55, 424-5 (1931)—A summary of articles by S. Nagas *et al* appearing in *J. Soc Chem Ind (Japan)* 33, (1930)
- Chemical composition—influence on manufacture and quality of cement. ALTON J. BLANK. *Rock Products* 34, No 2 62-3 (1931)—As the Al_2O_3/Fe_2O_3 ratio increases, kiln output decreases, fuel consumption increases, grinding becomes more difficult, quick setting and unsoundness tend to become more frequent. These conclusions are based on several years' observations in various plants. RAYMOND WILSON
- Sources of trouble in a cement plant. ALBERT HEISER. *Tonind-Ztg* 55, 433-4 (1931)—Correct sampling and accurate analyses are of primary importance. Ignorance and carelessness of employees as well as poorly working kilns must be considered. H. gives analysis of raw materials and shows how to calc. the correct mixt. for a product contg 76.4% $CaCO_3$ and a ratio $SiO_2, Al_2O_3 + Fe_2O_3 = 2.3$ B. A. SOTLE
- Transporting and elevating in cement works. A. C. DAVIS *Cement and Cement Manuf* 3, 631-40 (1930)—Belt, bucket, screw and pneumatic conveying systems are described. RAYMOND WILSON
- Power for cement works. A. C. DAVIS *Cement and Cement Manuf* 3, 925-45 (1930)—Elec. power equipment is described. RAYMOND WILSON
- New apparatus for measuring the consistency of cement pastes. KARL E. DORSCH

Zement 20, 26-8(1931)—A plunger is allowed to fall freely between guides upon the surface of the cement paste contained in a hard rubber ring. The paste consists of 300 g cement mixed for 3 min with about 28% H_2O . The depth of penetration of the plunger or the weight of the paste displaced may be used to compare consistencies. H. F. KRIGER

Indirect determination of lime in cement with sodium oxalate. A. HINER. **Zement 19, 1154-5(1930)**—A definite weight, 1.9832 g., of $Na_2C_2O_4$ equiv. to 0.8300 g CaO , is added to the soln of 1.000 g cement from which the SiO_2 and R_2O_3 have been removed. After several hrs' standing, the pptd CaC_2O_4 is filtered off and washed. The filtrate is made up to 500 cc., and the excess $Na_2C_2O_4$ remaining in soln is detd by titrating 100 cc portions with 0.10 N $KMnO_4$ soln in the usual way. H. F. KRIGER

Testing refractories for cement kilns. J. SCHAEFFER. **Zement 19, 1112-3(1930)**—A furnace is described in which cement or powd clinker is carried by the flame against the heated face of the refractory. H. F. KRIGER

Some observations and data in connection with the manufacture of portland cement from raw materials and fuels high in sulfur content. ALTON J. BLANK. **Rock Products 34, No. 8, 64-8(1931)** cf **C A 25, 143**—Clinkers of high SO_2 content yield cements which are unsound or show retrogression in strength. This effect may be due to CaS assocd with the $CaSO_4$. RAYMOND WILSON

The entropy of portland cement. GEOFFREY MARTIN. **Cement and Cement Manuf 3, 678-80(1930)**—In calcg the thermal balance of the rotary kiln it is necessary to consider the temp at which heat is available. A table of entropy changes for portland cement clinker is given. RAYMOND WILSON

Action of carbon dioxide and acetic acid on portland cement. HANS TH. BUCHERER AND J. W. MEIER. **Zement 19, 1134-8(1930)**—Various cements were titrated with 1% $AcOH$ in the presence of CO_2 in an attempt to dissolve out the CaO and MgO without the SiO_2 and R_2O_3 . The titration curves of the quantity of acid used against the time in seconds before the color change of the indicators occurred showed breaks or stages for most of the cements, indicating the soln periods of definite compds. SiO_2 and R_2O_3 were found in the soln in increasing amts as the CaO dissolved. The titration curve of lime under similar conditions showed no break. H. F. KRIGER

Chemical action of gypsum and other set-retarders upon portland cement clinker. L. FORSBERG. **Zement 19, 1130-4, 1155-60(1930)**—The sulfate ion disappears from 100 cc of a satd soln of gypsum when shaken in contact with 50 g of ground clinker. This is considered to be taken up by the Ca aluminates, since considerable Al_2O_3 dissolves in H_2O unless gypsum is present. When aq filtrates from clinker and gypsum suspensions were combined, 2 types of crystals were found, one spherulitic and the other in long needles. These crystals were found to consist of $3CaO \cdot Al_2O_3$ plus $xCaSO_4$, where $x = 2.38-2.75$. From 1 to 2% of $Ca(NO_3)_2$, $CaCl_2$, $CaBr_2$, $Ca(ClO_3)_2$ and $Ca(NO_2)_2$ was found to increase the setting time of clinker flour 2-3 hrs., while some sparingly sol salts such as $CaCrO_4$, $CaSO_3$ and CaS had little effect. The action of several Na , K and NH_4 salts as set retarders is also discussed. Complex compds of Al_2O_3 , CaO and $CaSO_4$ with strong hydration and crystn properties are thought to form first, serving as nuclei for the setting of the cement proper. H. F. K.

Free lime in portland cement clinker and soundness of cement. KATSUGO KORYANAGI. **Rock Products 34, No. 7, 74-5(1931)**.—See **C A 25, 1353**. R. W.

A tentative modification of the free-lime method [for portland cement]. H. R. BRANDENBURG. **Rock Products 34, No. 6, 68-9(1931)**—The addn of 1 g anhydrous $BaCl_2$ greatly hastens the soln of CaO in the alcohol glycerol mixt. R. W.

Crushing and preliminary grinding in portland cement manufacture. A. C. DAVIS. **Cement and Cement Manuf 3, 513-22(1930)**—Jaw, gyratory, roll, hammer and disk crushers and ball mills are described. Power requirements and characteristic size-reduction ratios are given for the various mills. RAYMOND WILSON

Grinding of cement clinker. A. C. DAVIS. **Cement and Cement Manuf 3, 1181-95(1930)**—Jaw and cone crushers and ball and tube mills are discussed. R. W.

Coal grinding. A. C. DAVIS. **Cement and Cement Manuf 3, 1050-72(1930)**—The coal grinding equipment used in the cement industry, with particular reference to unit systems, is described. RAYMOND WILSON

Quick-setting silicate of soda cements for acid-proof tank and tower construction. FOSTER DEE SNELL and HOWARD FARRAS. **Ind Eng Chem 23, 525-9(1931)**—Acid-proof cements composed of a mixt of inert aggregate and highly siliceous Na silicate set by evapn of water and drying. They are workable until they dry and give a weak bond until thoroughly dry. When an acidic material such as Na silosilicate is incorporated, the cements are workable for only a few min., set without drying and yield

a hard, brittle cement with a good bond. Alk. types contg Al hydrate and neutral types contg a mixt of Al hydrate and Na fluosilicate are intermediate in properties. They set after a few min to give a rubbery mass which sets further by drying. The neutral or alk. types are therefore best able to take up strains incidental to further building operations. Details of representative types of acid proof construction are given for which the original must be consulted.

FOSTER DEER SNELL
Tricalcium silicate. LANST JÄNECKE *Zement* 20, 26(1931), cf. *C. A.* 24, 1309—Further notes are given which deal with the controversy over the existence of the compds $3\text{CaO} \cdot \text{SiO}_2$ and $3\text{CaO} \cdot \text{Al}_2\text{O}_3$ or $8\text{CaO} \cdot \text{Al}_2\text{O}_3 \cdot 2\text{SiO}_2$ in portland cement. H. F. K.

Field determination of cement content of paving concrete. H. ALLEN *Eng. News-Record* 106, 513(1931)—A field method is described for detg the wt. per cu ft. of concrete as placed, from which data the yield and cement content per unit vol can be calcd for jobs on which proportioning is done by wt and the mixing water is accurately measured.

R. P. THOMPSON
Artificial control of the strength of concrete. III. C. R. PLATZMANN *Zement* 19, 1139-41(1930)—The use of certain hardening and acid proofing materials for concrete floors is discussed.

H. F. KATZ
Cinders as concrete aggregate. EINEA CHRISTENSEN, *J. Am. Concrete Inst.* 2, 583-646(1931)—Report of Comm. 203 of the Institute.

RAYMOND WILSON
Ideally graded aggregates and the quality of concrete. PAUL HEATEL *Zement* 19, 1087-90, 1117-21, 1141-4, 1167-9(1930)—The combined aggregate gradings designed by Fuller and by Otzen (slightly heavier in intermediate sizes than Fuller's grading) are compared in concrete with natural gradings. Both gravel and crushed basalt were used. Specimens of const cement content (300 kg per cu m) were made up to similar consistency and of nearly equal w/c ratio for the usual strength tests and for water imperviousness. The natural sand and gravel grading which was excessively fine showed by far the lowest strength and water imperviousness. The aggregates graded according to Fuller's curve gave concrete of excellent transverse and compressive strength for both crushed and rounded coarse aggregates, though slightly lower than those of Otzen's gradings. Greater water imperviousness resulted from the use of Fuller's grading than from any other. The mere presence of a fairly high cement content is not sufficient to insure strength or quality in concrete.

H. F. KRIZOS
Strength and resistance to sulfate waters of concrete cured in water vapor at temperatures between 100 and 350°F. DALTON G. MILLER *Proc. Am. Soc. Testing Materials* 30, Pt. 1f, 636-53(1930)—With some exceptions the trend of those cements that lost strength most rapidly in storage was toward greatest resistance to the action of sulfate waters. Curing in water vapor between 100° and 180° F did not generally increase the resistance of concrete to the action of sulfate water, in some cases a decrease was indicated. Curing in water vapor at 212° F and above markedly increased the resistance with increase of curing temps between 212° and 283° F for a 12 hr curing period. Curing at 212° F has been more effective in developing resistance when continued for 6 days than for 2. Absorption of concrete cured in water vapor at high temps is not a criterion of resistance to sulfate waters.

H. C. PARISH
Heat balance of the gypsum plaster kettle process. WALLACE C. RIDDELL *Rock Products* 34, No. 9, 66(1931)—A report on a 10-day test run showing 61% of total heat input utilized in calcination, 20% loss in stack gases, and 19% lost by radiation and otherwise.

RAYMOND WILSON
Roumanian trasses, their reactions and technical properties. A. STEPORE *Tonind-Zig* 55, 436-8(1931)—Exptl data are given to show that the baryta no method for detn of quality is valueless. S's method (*C. A.* 23, 686) is reliable. B. A. SOULE.

Scales for field tests of materials. HAROLD ALLEN *Eng. News-Record* 106, 533(1931)—A brief illustrated description of scales suitable for use in making sieve analyses, detg wt per cu ft and moisture content of aggregates, etc., for small projects for which the expense of a lab would be unwarranted.

R. E. THOMPSON
The active material of wood-conservation agents. E. PANNEWITZ *Z. Desinfekt u. Gesundh.* 23, 57-68(1931)—The inorg and org chemicals, in use or proposed, fungicides or insecticides in wood conservation are completely reviewed. Special consideration is given to German patent literature.

B. C. BRUNSTETTER
Hygroscopicity of woods after different kinds of drying. CARL G. SCHWALBE and KURT BERNDT *Kolloid Z.* 54, 314-26(1931)—Small rectangular blocks of pine (*Pinus excelsa* and *Pinus silvestris*) were subjected to 92% humidity after a drying treatment. Drying at 50° in a vacuum for 24 hrs and drying at 105° for 48 hrs decreased the water absorption. No difference was found on drying the wood in an O₂ stream or a N₂ stream. Preswelling the wood in water and then placing in the 92% humid air

increased the water absorption. The processes are irreversible only to a slight degree. Drying with CaH_2OH or CaH_2 does not change the water absorption. The swelling capacity may be detd by the measurement of the water absorption. A. F.

Apparent specific gravity apparatus [for stone, slag, gravel, etc.] (SHURF) 1. Accurate air separator for fine powders (ROLLFR) 1. Preparation of active MgCO_3 from magnesite (MATHERS, EAGLESON) 18. Distilling shale, coal, etc [to form pitch-like residue suitable for roads, roofing, etc.] (Brit. pat. 379,903) 22.

BUSSMEYER, HANS. *Wärmewirtschaft in der zementindustrie*. Dresden. T. Steinkopff. About 100 pp. About M 7, bound, about M 8.20.

Cement. CHARLES E. KRAUS. Ger. 520,033, June 18, 1926. See Brit. 253,929 (C. A. 21, 2778).

Cement. STEFAN KOHUT. Ger. 520,211, Oct. 9, 1929. In the manuf. of cement by a fusion process in which a mixt. of raw material and solid fuel is supplied to a shaft furnace supplementary pulverulent, liquid or gaseous fuel is supplied to the furnace, at a point above the tuyères, in order to reduce the production of CO in the furnace. The known methods of supplying supplementary fuel when smelting Fe in cupola furnaces may be adopted in the process.

Method and apparatus for grinding cement while cooling. CARL PONTOPPIDAN. Ger. 519,308, Sept. 23, 1928. Corresponds to Brit. 281,091.

Rotary kiln and cooling drum for use in cement manufacture. FRITZ LUTHER (to the firm G. Polysius). U. S. 1,797,831, March 24. Structural features.

Shaft and rotary furnaces in series for cement manufacturing. G. POLYSIUS A.-G. Ger. 516,526, May 3, 1928. Details of arrangement and interconnection are given.

Rotary-tube furnace and cooler for cement, ores, etc. G. POLYSIUS A.-G. Ger. 520,212, Sept. 3, 1926.

Apparatus for adding flux, etc., to cement mixtures. G. POLYSIUS A.-G. Ger. 517,652, Feb. 21, 1929.

Sample plates or tiles of colored cement. PHILIPP SCHUBERT. Ger. 517,700, Jan. 23, 1930. A layer of water-absorbing material (e. g., gypsum, felt, pumice, etc.) is formed on a ground plate of metal, etc., dried and the colored cement layer is then added.

Artificial aggregate for mortars and concretes. JOHN E. GREENAWALT. Fr. 606,067-8, May 24, 1930. See U. S. 1,780,713 (C. A. 25, 793).

Porous concrete. ENAMUELE STRACQUADAINI. Fr. 607,127, June 6, 1930. A cellular or porous concrete is made by adding powd. Al, Zn or Ca and a resin such as colophony to a cement, e. g., portland cement.

Concrete bricks, pipes, etc. G. R. TUSTIAN. Brit. 340,112, Dec. 3, 1929. Bricks, pipes and other articles of concrete are waterproofed and hardened by coating them, while still moist, with a compn. of cement, sand, powd. mica and water, subjecting them to steam satn. in a closed chamber, and drying.

Mixtures of concrete and asphalt. HERMANN MILKE and JOSEF OBERBACH. Fr. 695,265, May 7, 1930. Mixts. of concrete and asphalt are prepd. by adding to mixts. of ores and natural calcareous asphalt, mixts. of coal tar and artificial asphalt.

Macadam, etc. I. G. FARBENIND. A.-G. Fr. 697,286, June 12, 1930. A slag suitable for making macadam, paving blocks, etc., is made by adding to a slag from P. furnaces, the compn. of which corresponds approx. to Ca metasilicate, about 5-10% of Al_2O_3 . The Al_2O_3 may be added to the furnace charge in the form of Al_2O_3 or in the form of materials contg. Al_2O_3 .

Paving slabs. THE DUNLOP RUBBER CO., LTD. Fr. 696,044, May 23, 1930. Rubber paving slabs have the upper layer of relatively elastic rubber; the intermediate layer is of hard rubber of the consistency of ebonite, and both are seated in bitumen and bituminous aggregates.

Impregnating natural or artificial stone with bituminous material. KLEINBERGER & CO. A.-G. Ger. 519,129, May 9, 1926. The stone is heated to about 250° and then immersed in a bitumen bath having a lower temp.

Fireproof mortar. RHEINISCH & CO., VEREINIGTE KIESELGUHR- UND KOKSTEINGES. Ger. 517,959, Feb. 22, 1930. A fireproof mortar contains a mixt. of carbonates and (or) oxides of alk. earth metals and heavy metals, with an addn. of at least 2% of heavy metal oxide. In the example, the mortar contains Fe_2O_3 , MnO_2 , CaCO_3 , CaO calcined kieselguhr, powd. asbestos, gypsum and FeSO_4 .

"Artificial marble." F. RASES-GALCERAN. Brit. 339,674, Sept. 10, 1929. A

polished plate of material such as glass, "crystal," steel or brass which has been treated with oil and then chainail is placed in the bottom of a mold, and over this is placed three successive layers of material (1) a wet mixt of cement and ochre or other earthen coloring matter to which Zn white is usually added, (2) a dry mixt of fine sand and cement, and (3) a wet mixt of rough sand and cement. The product may be repeatedly treated with boiled linseed oil.

Roofing material. W. M. SHAKESPEARE (to International Copperclad Co.) Brit. 340,168, March 1, 1929. Structural features of composite roofing which may be formed with asphalted felt and thin Cu (which may be in the form of an electrodeposited strip).

Securing linoleum and similar plastic masses to floors, etc. ARTIFEX CHEMISCH FABRIK G. M. B. H. Ger. 517,817, Oct. 13, 1929. Use is made of a layer of an uncongealed mixt of oil and CS_2 to which mineral material may be added. Lactice and tar oil may be added. The oils mentioned are, olive, castor, soy and maize.

"Sheathing lumber" from redwood bark. I. LTON R. DARLING (to Pacific Lumber Co.) U. S. 1,797,901, March 24. Redwood bark is treated for the sepn. of its fibers by boiling in a soln. of a salt such as $Al_2(SO_4)_3$, capable of forming colorless tannates and producing a product which is suitable for making sheets or boards.

Photographic reproductions of wood grain. ALASA GES. ZUR HERSTELLUNG KUNSTLICHER OBERFLÄCHEN. Brit. 340,131, Jan. 26, 1929. A smooth board, before being photographed, is treated with a hard wax or with a polishing stone, such as agate or like non-smearing material, to alter the w of the surface parts without affecting the underlying parts. Cf. C. A. 24, 560.

"Wood substitute." S. SCHLESS. Brit. 339,853, Jan. 27, 1930. A product which may be made in laminated form comprises a mixt of pulverulent CaO or $Ca(OH)_2$ with pulverulent silica and cellulose or other fibrous material, the mixt being moistened, allowed to stand, pressed, heated to about $100-110^\circ$, treated with CO_2 and finished if desired by oiling, impregnating and polishing.

Apparatus for drying wood. CHARLES GOODALL. Ger. 517,714, Sept. 15, 1929. O or O_2 is added to the drying air.

Plant for drying wood. FRITZ HAAS. Ger. 521,005, June 18, 1929.

Preserving wood. SOC. ANON. DES ÉTABLISSEMENTS ARMAND BRAUTMARTIN. Fr. 695,479, May 12, 1930. Wood to be preserved by antiseptics has its surface perforated with a large no. of small holes by an instrument covered with sharp projections.

Impregnating wood with salt solutions. IMPREVA HOLZIMPRÄGNIERUNG UND HOLZVERWERTUNG A. G. Ger. 521,110, Dec. 1, 1925. The wood is steamed, rapidly dried at a temp. up to 150° , and then immersed while hot in the cold salt soln., e. g., a soln. of $HgCl_2$. The wood may be kept hot after drying by steaming again. App. is described.

Preserving organic material. SERGE PISSAREV. Ger. 514,981, Sept. 3, 1926. Wood, paper, fiber, etc., are preserved by impregnating with metal salts of naphtheneic acid in alk. NH_3 soln. and fixing with metallic Ni , compds. of cellulose. Salts of Cu, Zn or Ni are preferred.

21—FUELS, GAS, TAR AND COKE

A. C. FIELDNER AND ALDEN R. EMERY

Notes on recent developments in fuel technology. R. WIGGINTON. *Fuel in Science & Practice* 9, 497-9. 34-7(1930).—Brief reviews are given of industrial use of pulverized fuel, Lesthoman oil shale, gasification of coke breeze, prevention of gas explosions, industrial use of coal gas, oil in Germany, burning a hole in fog.

Graphical methods of fuel control. JAS. CUNNINGHAM. *J. Inst. of Fuel* 4, 216-22 (1931).—C shows that much of the fuel data entered into plant records can be used graphically to increase fuel economy. D. A. REYNOLDS.

Manufacture of alcoholic motor fuel. K. R. DIETRICH. *Z. Spiritusind.* 53, 349 (1930).—General discussion of the subject. S. J.

The addition of alcohol and other substituents to motor fuels. E. PYHÄLÄ. *Acta Chem. Fennica* 3, 123-4(1930).—A review. S. A. KARJALA.

Molasses as a fuel. A. RODZIEWICZ. *Gazeta Cukrownicza* 68, 133-5(1930). *Listy Cukrowar* 49, No. 26, Rozhledy 24.—Low monetary values of molasses suggested the use of molasses as a combustible. The sugar factory at Brzele Kujawski mixed coal with 20 and later with 10% molasses. The mass became tacky in the furnace and was

difficult to stoke. The warm molasses drained off the coal into the ashes before combustion occurred. Greasy spot deposited throughout the shafts. For economical usage the furnace will have to be altered so that the molasses will enter the hot coals and a copious blast of air will have to be supplied to prevent the formation of soot.

FRANK MARRIUS

New method of briquet manufacture. T SZAFYKA *Tűzoltószek* 2, 21-2(1929) — The method of colloid briquets is described. S S DE FINÁLY

Hungarian lignites. SÁNDOR KÁNDLY *Tűzoltószek* 2, 41-2(1929). — Drying alone cannot produce lignites with more than 1000-4500 cal. A suitable product of 7,000 or more cal can be obtained by distg lignites. Hungarian lignites contain about 8% tars, 50% of which can be used on distn and used for various purposes. S S DE FINÁLY

Importance of air occlusions and discharge of air in the manufacture of [fuel] briquets. J LUKÁRY *Soproni M. K. Hírv. 1. sz. Közl. 1930, 4-11*. Solidity of briquets after cooling is due partly to the solidification of binding substances and partly to the decrease of inner pressure caused by air occlusions. Solidification is improved by conducting pressure heat from the presses. Decrease of inner pressure of occluded air is facilitated by warming the material to 400° before pressing. S S DE FINÁLY

Modern carbonization technic in the United States and [as compared to] the Ruhr. W ROUTON *Brennstoff-Chem* 12, 111-3, 128-9(1931) — Coking practice, by-product recovery, markets, efficiencies, and economies are contrasted. F W JUNG

Low-temperature carbonization. W J SKILLMAN *Trans. Inst. Mining Eng.* 79, 151-69(1931), cf. C. A. 24, 3620. Low temp carbonization is divided into three groups according to whether the retorts are externally heated, internally heated by hot gases from an external source, or internally heated by combustion of a portion of the material under treatment. On another basis, the processes are classified as (a) continuous and (b) intermittent. Lab-scale and industrial scale expts are described and the commercial development is indicated. Several illustrations are shown and data tabulated. W H HOYNTON

Low-temperature carbonization of Donetzky coal. A P SHAKHNO AND I H. RAPOROST *Brennstoff-Chem* 12, 121-2(1931), cf. C. A. 24, 2860. Data results, with the Al retort, are tabulated for a series of coals from the Donetz district representing all 6 types in the rank classification of S and R. Analytical data for the coals used, the coke and tar obtained, including proximate, ultimate and coloring analyses are also given in table form. F W JUNG

Coal in 1930. G. L. EVANS *Mining J. (London) Ann. Rev.* No. Jan. 24, 1931, 10. ARDEN H. FERRY

Analyses of Washington coals. *Min. Mines, Tech. Paper* 491 (1911). The coal fields of Washington. S H AUST *Ibid* 1-11. Methods of mining and preparation of coal. H. P. YANCY AND S H AUST *Ibid* 12-22. Production, distribution and use. O S. KIRKLAND AND JOSEPH DANIELS *Ibid* 23-31. Relation of mine samples to commercial shipments. N H. SNYDER AND L. N. PETER *Ibid* 35-8. Analyses of delivered coal. N H. SNYDER AND L. N. PETER *Ibid* 39-65. — About 225 averages of samples of delivered coal are given. Analyses of mine samples. A C. FINTONER, H. M. COOPER, AND F. D. OSGOOD, *Ibid* 66-101. — Over 570 analyses are tabulated. Description of mine samples. F D OSGOOD *Ibid* 102-09. ARDEN H. FERRY

The washability of Transvaal coals. J MENDELSSOHN *J. Chem. Met. Mining Soc. S. Africa* 31, 191-5(1931) — It was unable economically to wash 8 Transvaal coals below 8 10% ash. Analyses of 5 samples are given. ARDEN H. FERRY

Hydrogenation of South African coal. P. N. LAYMAN, *J. Chem. Met. Mining Soc. S. Africa* 31, 199-200(1931); cf. C. A. 24, 2959-60, 6133; 25, 301. A H B

The Charleston coalfield. W H A. FENNELL *New Zealand J. Sci. Tech.* 12, 211-56(1931) — The location, history, geology and phys. characteristics of Charleston, N. Z., coal are discussed. The proximate analyses of 22 samples of the coal are tabulated. The coal is generally non caking, has a high S content, and burns to a buff-colored ash. W H. HOYNTON

The hydrogen in coal. A. FECLER, G. H. KENYON AND A. MCCULLOCH, *Fuel in Science & Practice* 10, 4-15(1931) — Four coals of different ranks were chlorinated and distilled under identical conditions. Carbonization was effected in 100° stages up to 1000° and the products of each stage were examined. CH₄ was the only paraffin hydrocarbon found in the gaseous products in measurable quantities. No tar was formed. Practically all the Cl in the coal was evolved as HCl. This evolution was almost complete at 600° except with one coal where it continued to 1000°. D A RHYNDORF

The swelling of coal during coking. K. BAUM AND P. HENNER, *Glückauf* 66,

1497, 1538(1930) *Fuel in Science & Practice* 10, 51-61 —An abstract of recent literature bearing upon the hydrogenation of coal. D. G. SKINNER. *Fuel in Science & Practice* 10, 109-37(1931) —A review of scientific press and patent literature with 218 references. D. A. REYNOLDS

Development of hydrogenation of coal. SZ. HANKISS. *Tűzeléstech* [N. S.], 2, 17-22(1930) —A review of the theoretical and technical development, present state and possible future of coal hydrogenation. S. S. DE FINÁLY

Distribution of carbon, hydrogen, nitrogen, sulfur and oxygen in the products of hydrogenation of a Hungarian coal. IMRE MAKRAY. *Tűzeléstech* [N. S.], 2, 25-8(1930), cf. *C. A.* 25, 188 —The coal used contained H_2O 2.60, ash 7.30, total S 4.23, combustible S 3.62, C 63.01, H 5.42, N 1.04, O 16.71%. net calorific power 6024 cal. The oil obtained contained 53% of the original C, H 30, S 2.9, N 1.7 and O 15%. The gas produced contained 21.7% of original C, H 4.7, no S, N 5.6 and O 46%. The water obtained contained 0.2% of the original C content of coal, H 10, S 0.4, N 1.6 and O 34%. S. S. DE FINÁLY

A study of the Marshall-Bird test for determining the agglutinating value of coal. K. A. JOHNSON AND H. T. YANCEY. *Fuel in Science & Practice* 9, 517-21(1930). See *C. A.* 24, 3343. D. A. REYNOLDS

The spontaneous ignition of coals. KÁROLY GÁTHNER. *Matematik és Természettud. Értelen* 46, 378-406(in German 407-8)(1929) —Ignition points of Hungarian brown coals were detd. in air and in O_2 currents. Coal consisting of smaller particles ignited more readily. Ignition is accelerated by the presence of unsatd. org. compds. The presence of S, moisture and humic acid showed no influence on the ignition temp. Data in the literature on the spontaneous ignition point of coals have value only in case grain size of coal is stated. S. S. DE FINÁLY

Coking coal. Its nature and synthesis. BALARAM SEN. *Proc. 15th Indian Sci. Congr.* 1928, 292 —The coking value of coal does not depend on its ash, volatile matter or fixed C, but is due to a resinous substance in the coal (called resinsol by S.) which can be extd. with a suitable solvent. Coal treated with a suitable solvent loses its coking power to an extent depending on the degree of extn. Poor coking coals can be made to cake by the judicious addn. of pitch or other suitable bond. Increases of volatile matter in the coking coal results in the lowering of its coking value. E. J. C.

Coal-dust recovery at Toronto Station. W. S. JOHNSON. *Power* 73, 409-11(1931) —Bag filters are used. D. B. DILL

The connection between ash content and specific gravity of Hungarian brown coals. J. FINKEV. *Soproni M. K. Bány. Ird. Főiskola Bány. és Koh. Orsz. Közleményei* 1930, 158-65 —Expts. made with sep. fractions of washed coal showed a mathematical connection between sp. gr. and ash content. A brown coal of Dorog gave, e. g., the following data:

Fractions	% of unwashed coal	Ash content %	Sp. gr.
1	8.5	5.6	1.32
2	56.8	11.6	1.41
3	14.1	28.1	1.57
4	20.6	71.0	2.17

Sp. gr. was detd. with the volumoscope of Kühn within 2 min. Graphical examn. of the above data gave the equation $A = 82d - 104$, in which A is ash content and d sp. gr. of the coal. S. S. DE FINÁLY

The possibility of manufacturing illuminating gas from Hungarian brown coals. ZS. GALÓCSY. *Tűzeléstech* 2, 42-7(1929) —An address. S. S. DE FINÁLY

The Fleissner method of drying lignites. H. KLEIN. *Intern. Berge. Bergtech* 23, 377-82(1930), cf. *C. A.* 24, 4916 —The method is well adapted for improving low-grade fuels, especially brown coal. The method consists of preheating the coal, heating under pressure with live steam, release of pressure and airing. A plant operating with this method is described. The properties and economic possibilities of dried lignites are discussed. A. HIRSCH

Plant and operating costs with the low-pressure Mont Cenis process. W. F. SCHOLVIEN. *Chem. Met. Eng.* 38, 133-5(1931), cf. *C. A.* 25, 1971 —Detailed plant and operating costs based on the data of German plants show that anhyd. NH_3 may be produced in the U. S. for 2.25 cents per lb. by the low-pressure Mont Cenis process. L. W. T. CUMMINGS

Combustion of natural gas. T. J. ESS. *Blast Furnace and Steel Plant* 19, 547-50(1931), cf. *C. A.* 24, 5135, 25, 269. W. H. BOYNTON

Results of laboratory experiments on manufacture of gas from Hungarian coals. ISTVÁN JAKÓBY AND ERVIN BECKER. *Tűzeléstech* [N. S.], 2, 1-6 (1930) —Seven typical samples of various Hungarian coals were tested for distn. A large content of S (1-6%) makes purification of gas too expensive. Coal of Nagybátóny (41.90% C) gave 13.9 cu m. gas per 100 kg. coal on high-temp. carbonization, 72% coke remained, with a C content of 34 kg. Half of this was converted to water gas and mixed with the original distn. gas. This mixed gas contained CO_2 5.5, O_2 0.8, C_2H_2 0.5, CO 28.6, H_2 47.3, CH_4 12.5 and N_2 4.8%, sp. gr. was 0.610, net calorific power, 3587 cal. S. S. DE F.

Catalytic preparation of a gas rich in methane from a mixture of water gas and steam. J. C. GHOSH AND K. CHAKRAVARTY. *Proc. 15th Indian Sci. Congr.* 1928, 148-9 —The authors have previously shown that a gas mixt. contg. CO and H_2 1:1 by vol. can be made to yield a fuel gas contg. over 25% CH_4 if passed over suitable catalysts at temps. 350-450°. The efficiency and life of certain catalysts have now been very considerably increased by introducing steam in suitable proportions in the reacting gas mixture. The following reactions take place simultaneously on the catalyst surface: (1) $2\text{CO} \rightarrow \text{CO}_2 + \text{C}$, (2) $\text{C} + \text{H}_2\text{O} \rightarrow \text{CO} + \text{H}_2$, (3) $\text{CO} + 3\text{H}_2 \rightarrow \text{CH}_4 + \text{H}_2\text{O}$, (4) $2\text{CO} + 2\text{H}_2 \rightarrow \text{CH}_4 + \text{CO}_2$. It is seen from (1) and (2) that CO reacts with steam giving CO_2 and H_2 . The H_2 so produced increases the ratio of H_2 to CO in the reacting gases and prolongs the life of the catalyst by preventing deposition of C. Catalysts (not mentioned) have been found which so regulate the rates of the reaction given above that the resulting fuel gases, after removal of CO_2 , contain more than 50% CH_4 . E. J. C.

Experiments on a Goffin gas plant. A. ROMWALTER. *Soproni M. K. Bány. Erd. Fosskola Bány. és Koh. Oszt. Közleményei* 1930, 166-77 —By a special method light hydrocarbons were obtained instead of low-temp. tars from S-contg. Hungarian brown coals. Lignite coke and distn. gas were secondary products. The amount of light hydrocarbons obtained was larger than that obtained by hydrogenation of low temp. tars produced by the usual carbonization. Results of large-scale expts. will be published later. S. S. DE F.

The origin and decomposition of organic sulfur compounds under gas-making conditions with particular reference to the role of the carbon sulfur index. JOHN C. HOLTZ. *Fuel in Science & Practice* 10, 16-30 (1931) —Exptl. work on the cracking of a gas oil contg. 3.61% S showed that CS_2 may be formed if a temp. of 649° is exceeded. CS_2 was formed on cracking a S-free oil in the presence of a gas from the same oil to which H_2S was added. An important loss of S to products other than CS_2 was noted. N. contg. small quantities of H_2S was passed over sugar charcoal and heated to incandescence. After a short time H_2S was formed and its concn. slowly increased. After more gas had passed, CS_2 appeared as a reaction product. The residue contained S, and concns. proportional to the surfaces per unit wt. of charcoal exposed. The effect of hot surfaces of pumice or charcoal free from S in removing org. S was studied. As the deposition of C and S on these surfaces increased they became ineffective for removal. Conclusion. Org. S compds. result from the decompn. of a complex of S and C. This theory was tested under equal conditions by passing N. contg. H_2S over sugar charcoal and cooling the reaction products to prevent reversion. CS_2 and other org. compds. were formed. Other factors affecting the formation of gaseous org. S compds. studied were: temp.; concn. of H_2S ; the presence of H_2O , CO , O or H_2 . An increase in temp. or concn. of H_2S caused an increase of total org. S formed as well as that of CS_2 . CO or O caused an increase in the org. S other than CS_2 . H_2 caused a decrease in the concn. of org. S. H_2O had little effect on the concn. of total org. S but reduced that of CS_2 . D. A. REYNOLDS.

Organic bases for gas purification. R. R. BOTTOMS. *Ind. Eng. Chem.* 23, 501-4 (1931). —Certain amines and amino compds. have the property of absorbing CO_2 and H_2S to form unstable carbonates and hydrosulfides that decompose at about 50° to liberate the gases and release the free bases in their original forms. Among the N bases found to be suitable are benzylamine, diaminopropanol, diethanolamine, diethylaminocethanol, etc. Operating methods are outlined. H. L. OLIN.

The bearing of fluctuations in the price of coke and by-products on the cost of production of gas. M. BARASH. *Trans. 2nd World Power Conf.* (Berlin) 2, 42-62 (1930) —Comparative costs of gas production in continuous vertical retorts, intermittent vertical retorts and horizontal retorts are shown. The value of the coke represents > 70% of the cost of coal. The credit for tar is sufficient to pay for the capital and maintenance charges, the labor costs and a portion of the fuel costs. The total value of coke and by-products forms so large a proportion of the credit side that the net cost of gas is likely to be more dependent on fluctuations in the prices of these resi-

duals than on variations in capital charges, fluctuations in labor costs, etc. These factors far outweigh any economies from large units. Figures given show that each variation of 1 shilling in selling price of coke affects the cost of gas production 3-4%. Cost figures for low temp. carbonization are given and show that only under very favorable conditions will the process pay. Tar, road tar, coal tar pitch, creosote oils and coal tar preservatives, cracking of tar, benzene, $(\text{NH}_4)_2\text{SO}_4$ and steam from waste heat are discussed. Benzene extrn is not a paying proposition and NH_4 liquor is a source of loss.

ALDEN H. TERRY

Phenols in lignite tara. F. V. HESSERT. *Z. anorg. Chem.* 43, 771-4 (1920).—A tar ext. representative of the phenolic compds. was dissolved in NaOH , polyphenols being eliminated by air oxidation. The remaining phenols were liberated and fractionated. Fractions were dissolved in 50% KOH , at 50° , and powd. $\text{K}_2\text{S}_2\text{O}_8$ was added to form phenylsulfonic esters. After satn with CO_2 , inorg. salts were pptd. with excess of alc. The dried esters, dissolved in abs. alc., were reprecip. with Et_2O . Alkyl groups were oxidized with alk. 6% KMnO_4 , the soln. was filtered, acidified with HCl , and evapd. Phenolcarboxylic acids were taken up with Et_2O , and again dried. Salicylic acid was extrd. with CHCl_3 , and the remaining acids were sepd. by fractional crystn. from H_2O . It identified the 3 hydroxybenzoic acids, the 4,1,3- and 5,1,3- $\text{HOCC}_6\text{H}_3(\text{CO}_2\text{H})_2$, the 2,1,4- $\text{HOCC}_6\text{H}_3(\text{CO}_2\text{H})_2$, the 3,1,2- and 4,1,2- $\text{HOCC}_6\text{H}_3(\text{CO}_2\text{H})_2$. From these, quant. and qual. conclusions regarding the original phenols were drawn. All the known coal-tar phenols were also found in the lignite tar. Considerable amts. of $\text{C}_6\text{H}_5\text{OH}$ were present in the xylol fraction. Trialkylphenols could not be detected. β -Substitution appeared to be the most favored.

K. H. ENGEL

An economic comparison between German and Japanese coking plants. S. AYABE. *J. Soc. Chem. Ind. (Japan)* 34, Suppl. binding 77 (1931).

E. H.

New methods of coke testing. WOLFGANG MELZER. *Arch. Eisenhüttenw.* 4, 225-35 (1930).—The phys. properties of coke samples from different locations in a retort differ greatly. The reactivity, structure, hardness and friability are influenced both by the type of coal and the method of coking. Coke produced in narrow retorts is not the equiv. of that from wide retorts. For blast furnaces which produce high-Si Fe a slow-burning, dense, hard coke is preferred.

H. F. JOHNSTONE

Reduction of ZnO by CH_4 or natural gas (DOERNER) 9. The reaction mechanism of C combustion at low pressures (SIMONEU) 2. Corrosion and metal protection in steam power plant equipment (STRUMPER) 9. Flame temperatures of hydrocarbon gases (JONES, *et al.*) 2. NaCNS as a gas drying agent (KELLER, NOBRT) 2. Apparatus and methods for precise fractional distillation—new method of gas analysis (POBIRZNIK) 2. Finished products from oil shale (BIRLOCH) 22. Butadiene hydrocarbons (Fr. pat. 696,705) 10. Fuel gas (Ger. pat. 529,281) 18. Dearsenification of acid [for manufacture of $(\text{NH}_4)_2\text{SO}_4$] (Fr. pat. 696,341) 18. Refining paraffin (Ger. pat. 517,744) 22. Distilling coal, etc. (Brit. pat. 339,963) 22.

FANSHAW, J. R. A Microscopical Study of Coal—Pennsylvania Anthracites and West Virginia Coking Coals. Lille G. SAUTAI 112 pp.

GUMZ, W. Feuerungstechnisches Rechnen. Leipzig O. Spamer 133 pp.

VI 8, bound, M 9

Some Features of the Glover-West System of Continuous Carbonisation in Vertical Retorts. Manchester, Eng. West's Gas Improvement Co., Ltd. 54 pp.

Fuel. KARL NEYNABEK. Ger. 516,762, Nov. 30, 1929. Details of a dehydrating press for crude peat, etc., are given.

Producing fuel for internal-combustion engines from water gas. MARCUS BAUTZKUS. Ger. 520,222, Feb. 21, 1925. See Brit. 252,785 (C. A. 21, 2551).

Motor fuels. ERNST G. F. MEYER. Ger. 520,011, June 13, 1929. Motor fuels consist exclusively of hydrocarbon distillates, Et_2O , and a readily volatile base, e. g., NH_3 , NH_4Me or NH_4Et , which diminishes the volatility of the Et_2O . The hydrocarbon distillate may boil within the range 70 – 200° , and may contain 1–5% b. up to 80° , 8–25% b. up to 100° and 20–65% b. up to 120° . The proportion of Et_2O may be 1–5%.

Motor fuels from plants or fruits containing fatty oils. GUTERHOFNUNGSHÜTTE OBERHAUSEN A. G. Brit. 340,107, Nov. 23, 1928. In a modification of the process described in Brit. 339,048 (C. A. 25, 2548), in which fatty oils are converted into products of low b. p. suitable for use as motor fuels by heating their vapors at high temps. with H_2 and preferably with catalysts, the plants or fruits contg. the oils in

their natural state are subjected to a similar treatment without preliminary extrn. of the oils. Soy beans yield 20% of an unsaponifiable oil b 60-200°. The catalysts may be used as aq solns of suitable salts and the hydrogenation may be effected in a rotary furnace.

Liquid hydrocarbons. PHYSICAL CHEMISTRY RESEARCH CO. Fr 696,610, June 3, 1930. Liquid hydrocarbons, particularly fuels, are obtained by bringing gases from the distn of solid fuels mixed with water gas under the action of an electrostatic field and under the action of a source of ultra violet rays. An app is described.

Fuel agglomerates. P. GLOESS. Brit 339,876, Feb 14, 1929. Fuel agglomerates formed with algine are impregnated with coal tar, wood tar, pitch or other org. coking material, followed by coking to render the agglomerates resistant to water.

Fire-lighting fuel. BRIQUETTING & CARBONISING SYNDICATE, LTD., and T. A. GOSKAR. Brit 339,985, Sept 17, 1929. Finely divided material such as coal or coke is mixed with comminuted peat and pugged and molded into briquets, and sawdust or paper pulp may be added and the peat may be treated with Na silicate before milling. The briquets are dipped in a bath of paraffin, paraffin oil and resin. Various details of manuf are described.

Briquet manufacture. W. BOOTH, V. BOOTH and W. J. BOOTH. Brit 339,868, Feb 6, 1930. Mech features.

Acetylene storage. AUTOGEN GASACCUMULATOR A. G. (to Gas Accumulator Co. (United Kingdom), Ltd.). Brit 339,899, May 30, 1929. Cylinders for holding C_2H_2 are filled with a cemented block of pumice, trass, natural or artificial slag or other material having the same coeffs of expansion by gas absorption and heat as cement.

Hydrogenation and desulfuration of hydrocarbons. PAUL L. F. NICOLARDOT. Fr 696,771, Aug 14, 1929. Gases from the distn of coal, oils, tars, etc., are transformed into substances resembling gasoline and petroleum, by subjecting the gases to a preliminary hydrogenation to sat the unsatd hydrocarbons and to render the sulfured products more sensitive to the final action of desulfuring materials, and then to a subsequent desulfuration and hydrogenation. Both operations are carried out at ordinary pressure or pressures not above 50 atm. The preliminary hydrogenation may be by means of nascent H produced by oxidizing CO with steam in the presence of a reacting mass at 350-500°.

Destructive hydrogenation. N. V. BATAAFSCHE PETROLEUM MAATSCHAPPIJ. Brit, 339,875, March 11, 1929. Co-pptd Fe and Al oxides (which may contain water and may be obtained by the action of NH_3 on a soln contg $FeCl_3$ and $AlCl_3$, washing the ppt. and drying it at 150° until it contains 10-12% H_2O) are used as catalysts in the destructive hydrogenation of materials such as coal, brown coal, tar, mineral oils and their derivs, wood, etc.

Filtering oils from destructive hydrogenation products, etc. I. G. FARBENIND A. G. Brit 339,681, Sept 13, 1929. Mixts of oil and carbonaceous matter such as produced in the destructive hydrogenation of coal, brown coal or tar or by refining oils with active C or with bleaching earth are filtered at an elevated temp. through a layer of powd or granular material or through a plurality of fine mesh filters, and the filter residue is freed from oil by distn at low temp., by extrn with solvents or by emulsification. The filtration may be effected in an inert gas, and numerous details and modifications of procedure are described, as is also an app.

Apparatus for distilling bituminous fuels. COMPAGNIE CONTINENTALE POUR LA FABRICATION DES COMPTEURS ET AUTRES APPAREILS and ALBERT BREISIG. Fr. 696,696, June 5, 1930.

Furnace for the low-temperature distillation or carbonization of coal, etc. CHARLES HONNAY. Fr 696,502, May 19, 1930.

Apparatus for carbonizing bituminous materials. HENRY LATDAM DOHERTY. Fr 696,414, Mar 15, 1930.

Apparatus for carbonizing bituminous materials. I. G. FARBENIND A. G. Fr 697,075, June 4, 1930.

Rotating furnace for carbonizing fuel at a low temperature. BRENNSTOFF-VER-SCHWELUNG G. M. B. H. Fr 697,085, June 4, 1930.

Electrically heated traveling-grate apparatus for low-temperature carbonization. TROCKNUNGS- VERSCHWELUNGS-, UND VERGASUNGS-G. M. B. H. Ger. 520,456, Feb. 3, 1928.

Agitating and conveying apparatus suitable for use in low-temperature fuel carbonization. MAXWELL MCGUINNESS. U. S. 1,793,026, March 24.

Low-temperature carbonization of coal. WALTER RUNGE (to International Coal Carbonization Co.) U. S. 1,797,796, March 24. Coal, while suspended in a "cloud-

like form, is subjected to the action of gases in a processing zone having an av. temp of approx. 575° , to effect deduction of a large portion of its volatile hydrocarbon content as gases and vapors, and the gases and vapors are cooled within the same app. to such an extent that a large portion of the vapors are condensed in a cooling zone with parts having a temp. as low as 150° and condensates thus obtained are returned to the processing zone where they are subjected to carbonization, gases thus obtained being sep'd. from the coke and removed and collected. App. is described. Cf. C. A. 24, 5137.

Directly heated distillation retort for coal, etc. GEORG MERKEL. Ger. 521,201, Jan. 16, 1929.

Rotary chamber oven for distilling coal, etc. HECTOR HARDY. Ger. 519,442, May 5, 1929.

Coking of soft coal. N. A. SILICA EN OVENBOUW MAATSCHAPPIJ. Fr. 696,385, June 2, 1930. Soft coal is coked by introducing it into the coke ovens in as divided a state as possible with appropriate charging means, the water content being kept between 6 and 10%.

Drying coal. BÖHMISCHE HANDELSGES. Ger. 520,369, Nov. 27, 1927. Coal is treated in a closed vessel with superheated water under pressure and when the coal has attained the temp. of the water the pressure is released. The water present in the coal itself is then liberated. The method is intended particularly for coarse lignite.

Centrifugal drying of peat and other materials. D. W. BERLIN. Brit. 349,156, Feb. 6, 1929. Peat, sawdust, waste wood, dry grasses or the like, which may be wet carbonized, is subjected to successive centrifugings, the first of which throws out the solid from assoc. liquid and the second of which (suitably utilizing app. with walls penetrable to liquid) throws liquid away from the solid. Various details of app. and procedure are described.

Gas. HUMPHREYS & GLASGOW, LTD., and ARTHUR R. CRIGGS. Fr. 696,082, May 24, 1930. In the manifold of gas by passing through coal the hot gases produced in the manifold of water gas from coke, the gases produced by injecting air are drawn off from the coke without passing through the coal and used for heating a recuperation chamber surrounding the retort containing the coal. The steam used for making the water gas is superheated by passage through this chamber. An app. is described.

Desulfurizing gases. I. G. FARBENIND. A.-G. Brit. 340,016, Sept. 16, 1929. S compds. are converted into H_2S and subsequently removed from coal gas, coke-oven gas, gases from cracking and distn. processes, etc., by use of catalysts comprising sulfides of elements of groups 2 to 7 of the periodic system together with metals of group 8 or heavy metals of group 1 or their compds. (preferably at a temp. of $200-400^{\circ}$). A suitable catalyst for treating coal gas may be prep'd. by pptg. Mo sulfide on Florida earth, applying Ni carbonate and reducing at 300° .

Washing gases. N. V. SILICA EN OVENBOUW Mij. Fr. 696,682, June 5, 1930. Substances contained in distn. gases, particularly C_2H_4 , are removed by bringing a solvent for the substance in a finely divided form into the gas, after which the gas charged with fog is brought to an elec. app. for purifying gases where the substance and solvent are ppt'd.

Removing nitric oxide from gases. GES. FÜR LINDE'S EISMASCHINEN A.-G. Ger. 521,031, Oct. 15, 1929. Traces of NO are removed from gases by washing them with solns. containing lower oxides of S or the alkali salts of the corresponding acids. Since the presence of free OH ions promotes the absorption, an alk. washing liquor is preferred, and the gases should be preliminarily freed from acid constituents other than CO_2 . The method is particularly useful in purifying coke-oven gases prior to sepg. constituents thereof by liquefaction, and may be put into effect in this case by adding an alkali sulfite to the alk. washing liquor used to remove CO_2 .

Apparatus for making fuel gas from crude liquid fuels. HEINO BALLRÖFER. Ger. 520,012, Feb. 8, 1929.

Washing coal gas, etc. GES. FÜR KOBLENTECHNIK M. B. H. Ger. 521,361, Dec. 13, 1929. A portion of the gas-washing water, which contains CO_2 and NH_3 , is with drawn from the still column at an intermediate point at which the water has lost most of its CO_2 content but still retains most of its NH_3 content. This portion of the water is returned to the gas washer.

Water gas. I. G. FARBENIND. A.-G. (Fritz Winkler and Paul Feiler, inventors). Ger. 516,655, Nov. 28, 1929. Addn. to 437,970. A method is described of producing water gas by conveying steam up a whirling column in which finely powd. fuel is burning. Cf. C. A. 25, 2573.

Apparatus for cooling gas such as carburated water gas. J. S. HARG (to Humphreys & Glasgow, Ltd.). Brit. 339,883, May 6, 1929. Structural features.

Plant for generating gas in furnaces. CLARK C. BOARDMAN (to Thermoatomic Carbon Co.) U S 1,797,256, March 21 A generator whose rate of discharge is subject to quick rises is assocd with a chamber contg primary cooling means such as a water spray cooling the gases discharged from the generator, with a second chamber for receiving the gases from the cooling chamber, and with auxiliary cooling means controlled in accord with the temp in the second chamber Various details of construction are described

Gas producer. J S HAUG (to Humphreys & Glasgow, Ltd) Brit. 339,743, Dec. 7, 1928 Various structural details are described relating to a water-jacketed water gas producer having a rotary grate with a central tuyère

Gas producer. ALDO LOMBARDI Fr 695,817, May 19, 1930 A device is described for making a gas rich in CO from wood

Gas producers. MASCHINEN- UND FAHRZEUG FABRIKEN ALFELD-DELLIGSEN A G Ger 517,590, Jan 16, 1927 The producer has an inclined grate for the introduction of air and a lower roller grate partly immersed in water

Producer for gasifying fuels. MATTHIAS FRÄNKL Ger 514,194, Dec. 3, 1929 A producer comprising two shafts communicating at the bottom is used Steam is passed in at the top of one shaft, and O or oxygenated air is passed in at the bottom of the shaft The gases are withdrawn from the top of the other shaft The functions of the shafts are interchanged from time to time

Recuperator for gas-heated gas producer. JEAN SAUVAGEOT and HENRY DIFTER LEV Fr 696,052, May 24, 1930

Cooling the gas liquor of gas producers. WILHELM OEFVERBERG Ger 517,977, Aug 9, 1924 Details of the app are described

Apparatus (with thermostatic control) for producing lampblack by combustion of natural gas, etc. BRADLEY W RUMARGER (to Imperial Oil & Gas Products Co.) U S 1,797,368, March 24 Structural features

Tars or pitch. COMPAGNIE DES MINES DE VICOIGNE, NOFUX ET DRACOURT Fr 696,245, Sept 7, 1929 Tars or pitch from the distn of coal lignites or schists at high or low temp are treated to eliminate oxygenated asphaltic or acid substances by pptn with appropriate solvents such as petroleum, gas oil, etc., or by fixation with alk. solns The purified products are hydrogenated

Acid treatment of wax-free tar. THERMAL INDUSTRIAL & CHEMICAL (T. I. C.) RESEARCH CO., LTD., and C. O. COYDRUP Brit 339,640, Aug 6, 1929 In connection with the agitation of wax free tar with acid such as H_2SO_4 at a temp of about 50–80°, the mixt of oils and tars (either with or without previous sepn and remixing) is neutralized, dehydrated, and heated until of the desired consistency, and the oils may be sepd and fractionated and the residue and oils (except those oils distg 325–395°) mixed with the tar after neutralization Various auxiliary steps and modifications of procedure are described

Filtering oleaginous or tarry liquids. I G FARBENIND A.-G Brit. 339,636, June 7, 1929 A filtering diaphragm such as that of a described rotary vacuum filter is prepd for filtering oleaginous or tarry liquids by applying to it a layer consisting of finely divided carbonaceous residues contg only, tarry or asphaltic substances such as those derived from destructive hydrogenation or heat pressure treatment of coal, peat, shale, tars or mineral oils (preferably after the residues have been heated to 250–300° to drive off lighter fractions)

Distillation apparatus for removing tar from waste waters from cokeries. PAUL GRUHL Fr 696,837, June 7, 1930

Coke. GUSTAV HILGER Fr 696,906, May 15, 1930 Canals for the evacuation of gas are formed in the charge of fuel used for making coke or semi-coke so that the gases are led from the hot walls of the furnace to the colder center of the charge

Coke. SOC GUSTAV WIPPERMANN, MASCHINENFABRIK, STAHLWERK UND EISENGIESSEREI G M B H Fr 696,647, June 4, 1930 Metallurgical coke is improved by adding coke dust of irregular-sized grain of dimensions below 3 mm obtained by crushing under pressure the coal used for making the coke

Method of dry-cooling semi-coke. N-V CARBO-UNION INDUSTRIE MAATSCHAPPIJ Ger 519,441, Oct. 5, 1924

Coke ovens. ARNOLD BECKERS Fr 695,632, May 15, 1930 Construction of tight fitting doors is described

Coke ovens, carbonization chambers, etc. WOODALL-DUCKHAM (1920), LTD., and ARTHUR McD DUCKHAM Fr 697,233, June 11, 1930

Horizontal coke oven. CARL OTTO U S 1,798,129, March 24,

- Vertical chamber oven for producing gas and coke. HEINRICH KOPFERS. Ger 521,105, Feb 2 1926
- Regenerative coke oven. HEINRICH KOPFERS A-G (Friedrich Totzek, inventor). Ger 520,073, Dec. 13, 1927
- Coke-oven battery with horizontal coke chambers. JOSEPH BECKER. Ger 516,532 Jan. 25, 1925
- Heating coke ovens. UNION CHIMIQUE BELGE, SOC ANON Belg 372,788, Sept 30 1930 The rate of feed of heating gas is maintained const., but in the periods between inversions the amts. of air introduced at various points are varied. No 372,789 provides for the use of a heating gas of variable compn., or consisting of a mixt. of 2 or more constituents, the proportions of which can be modified as required during the interval between 2 inversions of the heating medium
- Coke-oven-heating system. WOODALL-DUCKHAM (1920), LTD., and A. M. DUCKHAM Brit. 337,641, Aug 6, 1929 Structural features

22—PETROLEUM, LUBRICANTS, ASPHALT AND WOOD PRODUCTS

W. F. PARAGHER

Hydrogenation of petroleum oils. RALPH H. MCKER AND ANTONI SZAYNA *J Inst Petroleum Tech* 17, 121-33(1931)—The mechanism of hydrogenation is discussed from a consideration of the data in the literature L. W. T. CUMMINGS

Decolorization of petroleum with acid clay. L. K. YAMAMOTO AND H. ISHIKAWA. *J Soc Waseda Applied Chem* 1930, No. 12, 6-13, *Mém Faculty Sci Eng Waseda Univ* 1930, No. 7, 28—The power of acid clay to decolorize petroleum products increases with decreasing particle size II. III. *Rept Tech Chem Lab. Waseda Univ* 1930, No. 13, 5-7, 7-14 *Mém Faculty Sci Eng Waseda Univ* 1930, No. 7, 29, 30-1—The acidity and petroleum-decolorizing power of acid clays decrease as the temp. of treatment increases from 100° to 1000° HCl treatment increases and alkali treatment decreases the decolorizing power of acid clay. L. W. T. CUMMINGS

Synthesis of petroleum hydrocarbons from hydrogen and carbon monoxide at ordinary pressure. II. K. KOBAYASHI, K. YAMAMOTO AND H. ISHIKAWA. *Mém Faculty Sci Eng Waseda Univ* 1930, No. 7, 20-7 of C A 23, 5035—Nine catalysts composed of Co, MnO and Cu were used in the synthesis of oil from CO and H₂. Active C formed from the decompn. of the CO is considered the basic material for subsequent hydrogenation and polymerization to oils. CH₄ was formed simultaneously. L. W. T. CUMMINGS

Luminous stationary flames: the quantitative relationship between flame dimensions at the sooting point and chemical composition, with special reference to petroleum hydrocarbons. S. T. MINCHIN *J Inst Petroleum Tech* 17, 102-20(1931)—The tendency of hydrocarbons to smoke as detd. in a Weber photometer lamp was found to decrease in the naphthalene and benzene series as the no. of C atoms increased, but increased in the paraffin series. This property also increases in synthetic mixts. of paraffins and aromatics or naphthenes as the concn. of either of the latter is increased L. W. T. CUMMINGS

Cretaceous limestone as petroleum source rock in northwestern Venezuela. HOLLES D. HEDBERG *Bull Am Assoc Petroleum Geol* 15, 229-46(1931)—The lithology, microscopic character and occurrence of oil in La Luna and Cogollo limestones are compared in detail. The conclusion is that the oil in the former is indigenous, whereas it has migrated into the latter. The La Luna formation is probably a petroleum source rock. The specific source of the oil generating material in these sediments is discussed. ALDEN H. EMEY

East Hackberry salt dome, Cameron Parish, Louisiana. A. J. BAUERNSCHMIDT, JR. *Bull Am Assoc Petroleum Geol* 15, 247-56(1931)—The Hackberry oil is a typical heavy Gulf Coast crude oil ranging from 18° to 24° Bé (av. 22°) A. H. E.

Finished products from oil shale. WM. BLELOCH *S African Mining and Eng J.* 41, Pt. II, 393(1930)—The following products are obtained from Scottish shale oil gasoline (motor fuel), extg. spirit (for removal of oil from seeds and bones), solvent naphtha (mfg. rubber goods), 0.741 naphtha (safety lamps), 0.750 naphtha (flare lamps), white spirit (paints), solvent (removal of naphthalene from coal gas and manof. of linoleum), various grades of illuminating oil, gas oils, fuel oils, cleaning oil, batching oil (jute

spinning), lubricating oil, residual oil for compounding thick greases, paraffin wax, $(\text{NH}_4)_2\text{SO}_4$ and coke. ALDEN H. EMERY

Oil in 1930. ANON *Mining J* (London) *Annual Rev.* No. (Jan. 24, 1931) 10
The oil industry of the U. S. S. R. during 1930. W. LEONTIEV *Ibid* 11-2 —Oil production and consumption in Russia increased at a greater rate than ever before; production increased 17.1% in 1929 and 26% in 1930. Refinery production increased 37.1%. Cracking is almost negligible. ALDEN H. EMERY

The pyrogenic decomposition of paraffin oil by different catalysts in air and hydrogen. R. ODA *J. Soc. Chem. Ind. Japan* 34, Suppl. binding 58-62(1931) — Fe_2O_3 , Al_2O_3 , Nb_2O_5 , Fe , MgO , ZnO , SiO_2 and fullers' earth have been tried. SiO_2 , the best catalyst, gave 73% of cracked oil, 70% of which boiled between 35° at 760 mm. and 200° at 9 mm. The original oil b. 83-215°. V. F. HARRINGTON

Examination of paraffin products by means of the refractometer. W. J. PIOTROWSKI and J. WINKLER *Erdöl u. Teer* 6, 447-8, 463-5(1930). cf. *C. A.* 24, 233 —The method of Diggs and Buchler (*C. A.* 21, 516) is adapted to a rapid detn. of small percentages of oil in paraffin wax, by simplifying the isolation of the pure components. "Blue oil" (from Polish sources), mixed with amorphous SiO_2 (2:1), was chilled to -21°. A few drops of the pure oil were obtained by filtration through a Gooch crucible. Other samples were prep'd by dissolving paraffin products from various sources in Et_2O and reprecipitating paraffin by addn. of EtOH (1:1). All products gave closely agreeing values for n_D at 14.887°. It was found that oils were selectively removed from paraffin wax by fuller's earth (2:1) at 80°. A pure paraffin retained its n_D value after such a treatment, various com. samples and artificial mixts. of the components gave const. values for the pure paraffin after a single filtration (1.4370). The analysis of a mixt. requires only the detn. of its n_D value, if the values for the pure components have been det'd. Applied to known mixts., the method was accurate within 0.2%. Holden's method gave results which were higher by 1.1-2.6%. K. H. ENGEL

The chemical aspect of drilling muds. ALEXANDER DUCKHAM *J. Inst. Petroleum Tech.* 17, 153-62(1931) —Drilling mud is used to carry cuttings from the bit to the surface, to keep the cuttings in suspension, reducing settling when circulation ceases, to act as a lubricant for the drill pipe and between the formation and the bit, to minimize the abrasive action of sand and rock particles in the pumps, to keep gas, ml and water back in the formation, to plaster the walls of the hole and of cavities and exhausted sands, and to protect the operators against blowouts. Mud must often be "doctored" as to sp. gr., viscosity and colloidalty. Sp. gr. is raised by adding barites. Bentonite is used very generally to raise the colloidalty, though no standardization has been worked out. It has been suggested that 0.02% of bentonite should be added to the mud for every 1% of settling which would otherwise occur in 24 hrs. By adding quantities of Na_2CO_3 of the order of 0.2% wt. may be increased as desired without increasing the viscosity. Costs of doctoring are considered. EMMA E. CRANDALL

The reactions of olefins with sulfonic acid. W. R. ORMANDY and E. C. CRAVEN. *J. Inst. Petroleum Tech.* 17, 153-7(1931) —A series of olefins of C content C_4 , C_5 , C_{10} and C_{18} was treated with H_2SO_4 of concns. from 81 to 99.5%. The olefins of higher mol. wt. resisted the action of the H_2SO_4 better than those of lower mol. wt., when acid of 81.0 to 92.8% concn. was used. For the acids of highest concns. this was not true. The statement of Howes and Nash, that the olefins which are stable to KMnO_4 are also fairly resistant to H_2SO_4 , was not confirmed (cf. *C. A.* 24, 333, 3107). E. E. C.

Twenty-second semiannual motor gasoline survey. II. Additional data. A. J. KRAEMER and E. C. LANE. *Bur. Mines, Rept. of Investigations* 3092, 33 pp (1931), cf. *C. A.* 25, 806 —Results of the detn. of Cu strip corrosion, color, S content by the lamp method, Cu-dish evapn. residue, gum content (steam-oven method) and anti-knock value are given. Only 1 of the 298 samples failed to pass the Cu strip test; 23.7% of the competitive-priced gasolines and 54.6% of the premium-priced motor fuels were artificially colored, for the others, color was without significance, and 85.6% of the samples had < 0.10% S (lamp method). There is no relation between doctor test and color S or gum content. The av. octane no. of the competitive gasolines is 59 (51-68) and of premium motor fuels 74 (65-80). ALDEN H. EMERY

Gravity of natural gasoline and percent evaporated at 140° F. are related. S. S. SMITH. *Natl. Petroleum News* 23, No. 4, 31-2(1931) —From data obtained from inspection of a large no. of Mid Continent gasolines it is shown that gasolines with 70% evapd. at 140° F. have gravities of from 85 to 89 and when they are controlled to 26 lb. vapor pressure they will fall close to 70% evapd. at 140° F. and have gravities of 86 ± 2. R. W. KELLY

Gasoline dopes. H. C. DICKINSON. *Ind. Eng. Chem.* 23, 517-9(1931) —Based

has a Krämer Sarnov m p of 37°, no phenol reactions, free asphaltic acids 1.37%, lactones and anhydrides 2.20%, asphaltenes 2.11%, asphaltic resins 41.29% and only constituents 51.25%. Elementary analysis gives C 81.77, H 10.85, S 1.29, N 0.94, O 4.07 and ash 1.08%.

FRANK MARESH

The Ragusa asphalt rocks. ANDREA LA PORTA *Ind. mineraria* 5, 65-77 (1931) —A ton of Ragusa bituminous limestone contains 40-65 % of org combustible matter of higher calorific value than coal. At 450° it loses 1/2 of the hydrocarbons. A rock with 4% asphalt yields 2.5% of mineral oil. The bitumen contains 84.9% C, 11.7% H and 3.4% S. The consts of the raw mineral oil are d_{44} 0.960-0.965, viscosity (Engler degrees) at 20°, 20,000, at 50°, 3,500 and at 100°, 1,300. Pensky Martens flash 126°, fire point 158°, ash 0.008%, hard asphalt 1%, soft asphalt 0.15%, 32% remains unchanged by H_2SO_4 . S 2.19% org acidity (oleic acid) 0.35%, org acidity (SO_4) 0.049%, no mineral acidity, no alkalinity fractions under 300° 18-22.3%, Distn at Strasburg gave under 225° 1%, 225-300° 13%, 300-350° 35%, 350-380° 27% and a residue of 24%. Distn in a vacuum at 50 mm, at 200-200°, gave 28% of oil with a Pensky Martens flash of 158°, viscosity (L 35°) 1.5, from 200° to 345° there was 30% of oil with a flash point of 220° viscosity (F 35°) 41.8, at a pressure of above 50-100 mm above 345° there was 9% of oil with a flash point of 98°, viscosity (E 35°) 8.4 residue 6.6% loss 2.5%, 1.80-3.10% paraffin. The oil has cyclic, naphthenic and other unsatd compds, org S compds, little H, a high sp gr, very little hard asphalt and no resinifiable substances. The presence of unsatd compds makes it suitable for producing lubricants. Cracking yields 50% of light compds.

R. S.

Atjeh turpentine, its present quality and properties. IR. W. SROOV *Ber. Afdel. Handelsmuseum Ver. Koloniën Inst.* No 57, 30 pp (1930) —A description of the tapping and distn app used in Sumatra and tables of the principal consts of the turpentine are given. By the use of modern app an improved product is obtained, which, however, cannot as yet replace the superior product imported from the United States and France. The reason for this condition is discussed in detail.

J. C. JERRINS

The iodine number of turpentine oils. V. KUBELKA AND S. ŽURAVLEV *Chem. Listy* 25, 124-8 (1931) —Using the Hanuš, Margosches and Rosenmund methods for detg the 1 no of a French turpentine, K. finds the action proceeding in 2 phases: (1) a rapid and active phase lasting 3 hrs in the M and R methods and 30 min in the H. method, (2) a slow and prolonged phase due to side reactions. The wt of the sample alters the 1 no, it should not exceed 0.1 g. The 3 methods gave the following 1 nos: H 190-200, M 280 and R 300. If the turpentine is pinene with 1 unsatd bond, 1 mol of pinene adds 2 equivalents of I and demands an 1 no in the range 185-200, this is the value given by the H method. If 1 mol of pinene combines with 4 equivalents of I, the 1 no must be 370, this value is approached by the R. method. In the first phase of the reaction, the pinene unsatd bond combines with I, later, in the second phase, the internal chain is attacked and gives rise to free bonds which attach I. The Hanuš method gives results which have a sound theoretical foundation. Recommendations had been given that distn products of turpentine and pine oils could be differentiated by the 1 no. By the H method on 0.1-g samples and with the time of reaction limited to 30 min, 15 turpentine oils of standard grades were analyzed. The 1 nos ranged from 154 to 202. The 1 no cannot be used for differentiating fractions of pine oils, etc. Changes in 1 nos of oils during 15 yrs were insignificant.

FRANK MARESH

Genetic connection between salt deposits and petroleum (WEINFELD) 8. Anomalous velocity distribution in thin lubricant films (UMSTÄTER) 2. Compaction of sediments (TRASK) 8. Apparatus and methods for precise fractional distillation (PODNIELNIAK) 2. Conditioning refinery water (BORN) 14. Butadiene hydrocarbons (Fr. pat. 696,706) 10. Hydrogenation and desulfuration of hydrocarbons (Fr. pat. 696,771) 21. Solvents and emulsifying agents [for mineral oils] (Fr. pat. 697,102) 27. Apparatus for extracting oil from bleaching earths (Ger. pat. 520,169) 27. Destructive hydrogenation (Brit. pat. 339,875) 21. Filtering oleaginous liquids (Brit. pat. 339,636) 21. Filtering oils from destructive hydrogenation products, etc (Brit. pat. 339,681) 21. Electrically heated apparatus for cracking oils (Brit. pat. 339,649) 4. Desulfurizing gases [from cracking and distillation] (Brit. pat. 340,016) 21. Absorbing olefins with H_2SO_4 (Brit. pat. 340,098) 10.

Petroleum. YVES CORNIC. Fr. 696,163, May 27, 1930. Petroleum is obtained by treating CH_4 with nascent CO_2 ; the O of the CO_2 burns a part of the H of the CH_4 , which then condenses to petroleum. The CO_2 is obtained from carbonates dissociated.

by heat with or without the addn of substances favoring dissocn and with or without pressure

Cracking petroleum oils. CORNELIUS B. WATSON (to Gyro Process Co.). U. S. 1,707,305, March 24. Petroleum, from which free gasoline has been removed, is vaporized in a still and the vapor is conducted to a cracking app and cracked by heating; highly heated cracked vapors are passed to a sep receptacle through a closed coil within which charging stock is passed to effect vaporization of its low b p constituents, and the high b p portion of the charging stock is discharged into the receptacle in liquid form for admixt with vapors discharged from the cracking app., to lower the temp of the cracked vapors and arrest further cracking. An arrangement of app is described.

Cracking hydrocarbons. MARTIN B. SCHUSTER. Ger 521,032, Jan 18, 1927. See Brit. 289,556 (C. A. 23, 696).

Pressure cracking of liquid hydrocarbons. SAMUEL J. M. AULD and PERCY IF. HERRING (to Anglo-Persian Oil Co., Ltd.). U. S. 1,798,034, March 24. The liquid is heated in a single stream to below a detd limit of the eventual temp of cracking, then passed in narrow streams at low velocity in a direct and relatively short course into a reactor, and subjected in the narrow slowly moving streams having a very low cross-sectional temp gradient to heat uniformly and gradually applied to a cracking temp without causing consummation of the cracking reaction, and the liquid is maintained in the reactor to consummate the cracking before withdrawal of the treated liquid. App is described.

Cracking hydrocarbon oils. F. P. E. S. DUTLAN. Brit. 339,841, Jan 21, 1929. Cracking is effected by heating oils, under pressure sufficient to maintain them liquid, to a temp somewhat below 360°, while flowing through a tubular heating furnace at a speed of 6 to 72 cm. per sec for a tube of 10 mm diam, so that the oils are heated for about 1.0-1.5 min. only. Various details of app are described.

Cracking oils. PETROLEUM CONVERSION CORP. Brit. 340,030, Sept. 19, 1929. Oil vapors are brought into contact with a heated carrier gas supplied in stages so as to maintain a substantially const. temp level throughout the operation. App and details of temps and other features of procedure are described. Cf C. A. 25, 410.

Cracking oils. A. P. SACHS and E. W. BEARDSLEY (to Petroleum Conversion Corp.). Brit. 340,021, Oct. 6, 1928. In a vapor phase cracking process in which heat for the cracking is supplied by a carrier gas the oil is fractionally vaporized and the sep fractions are added in stages to the carrier gas so that the lighter fractions are subjected to a longer period of conversion. An arrangement of app and various details of operation are described.

Cracking and hydrogenating oils, etc., in the interior of a compressor. MARCUS BRUTZKUS. Ger 520,223, Feb. 21, 1925. See Brit. 252,787, (C. A. 21, 2556).

Coatings for oil-cracking retorts. ALBERT C. HOLZAPFEL. Ger 519,085, July 19, 1927. See Brit. 291,585 (C. A. 23, 1260).

Refining hydrocarbons in vapor phase. FRANK A. AFGAR (to Sinclair Refining Co.). U. S. 1,797,255, March 24. The vapors are passed through fuller's earth or other suitable adsorptive catalyst in 2 different bodies through which the vapors are successively passed, and the sequence of passage of the vapors is periodically reversed before complete deactivation of the first body of adsorbent occurs. An arrangement of app is described. Cf C. A. 24, 718.

Refining hydrocarbons in vapor phase. EUGENE C. HERTHEL (to Sinclair Refining Co.). U. S. 1,797,262, March 24. Hydrocarbon vapors such as those from cracking processes are passed through an adsorptive catalyst such as fuller's earth and high b p polymers produced are then sep'd and the vapors are subjected to fractional condensation to sep substances of higher b p than desired in the final product. Oil components are vaporized from the polymers and higher b p condensate and are condensed free from tar including the high b p polymers and this oil condensate is passed through the adsorptive catalyst with the hydrocarbon vapors to be refined during the refining operation, as a washing agent. An arrangement of app is described.

Refining hydrocarbon oils. IRVING C. CARPENTER and ARTHUR R. MOORMAN (to Contact Filtration Co.). U. S. 1,797,715, March 24. A body of oil such as crude petroleum is maintained at a temp above the vaporization temp of water but not substantially above the vaporization point of a desired fraction of the oil, by heating a stream of oil withdrawn from the body to such intermediate temp., sep'g vapors from the heated liquid and returning a portion of the heated liquid to the body of oil to maintain a substantially const. volume, with sep feeding onto the surface of the body of a stream of oil to be treated and sep removal of vapors from the body. App is described. Cf C. A. 24, 3638.

Refining oils. THE SHARPLES SPECIALTY CO. Fr. 696,461, May 9, 1930. Oil of low m. p. is obtained in a continuous manner from mineral oil contg. paraffin by submitting the mineral oil to a continuous vaporizing heat and obtaining from the resulting mixt. of vapors and liquid by a continuous fractionated condensation with reflux, a fraction contg. paraffin which ppts. during a lowering of temp. of this fraction in a form in which it may be removed by a continuous centrifuging process.

Hydrocarbons of low boiling point. I. G. FARBERND. A-G. Fr. 695,496, May 13, 1930. Hydrocarbons of high b. p., mineral oils, distn. residues, etc., are preliminarily treated in the liquid state at high temps. with H₂ under pressure, in the presence of catalysts contg. elements of group 2 to 8. The S and O compds. contained in the starting materials are destroyed, to a great extent, without any appreciable amt. of decompn. or hydrogenation of the hydrocarbons. Afterward the purified materials freed from the catalyst, are decomposed by heat in the presence of H₂. Cl. C. A. 24, 5470; 25, 1536.

Distilling shale, coal, etc. F. ESLING. Brit. 339,963, Sept. 12, 1929. Bituminous shales, cannel coals, torbanites, etc., are distd. in admixt. with petroleum or its derivs. at ordinary, reduced or slightly increased pressure, to form distillates of low S content and an un-coked pitch like residue suitable for road surfacing, roofing, etc. Vegetable oils such as linseed, cottonseed, rape, olive or tung oil may be added, and the shale is preferably powdered and may be mixed with lime to assist in removal of S. Steam may be injected during the distn.

Column still suitable for use in fractionation of oil vapors, etc. J. V. GAUDET and A. ABRAMSON. Brit. 339,663, Aug. 8, 1929. Various structural details are described.

Oil-refining residues. I. G. FARBERND. A-G (Richard Michel, inventor). Ger. 516,653, July 27, 1928. The ext. left by the refining of oils by liquid SO₂ is worked up by treating with olefins in the presence of catalysts. Thus, such an ext. of viscosity 1.4°E. at 20° is treated with C₂H₄ at 20 atm. at 140° in the presence of AlCl₃ until no further C₂H₄ is absorbed. The product is washed with water and distd., giving a yellow oil of viscosity 4°E. at 20°, and a resinous residue. A further example is given.

Refining paraffin. A. RIESBACH and MONTANWERKE A-G. Ger. 517,744, April 5, 1929. Paraffin is deprived of smelling and tasting impurities by treating it with alkali in solid form or in coned aq. soln. or suspension. The paraffin is then sep'd. without distn. Adsorbants may also be added. The process may take place under pressure and is particularly suitable for lignite tar paraffin.

Filter for gasoline, etc. ANTOVIN BOULADE (to Soc. du carburateur Zenith). U. S. 1,797,399, March 24. Structural features.

Apparatus for refining "cracked gasoline" by use of fuller's earth, etc. RUDOLPH C. OSTERSTROM and ROBERT T. TUCKER (to Pure Oil Co.). U. S. 1,797,513, March 24. Structural features.

Liquid fuels. GENERAL TECHNICAL CO., LTD. Fr. 696,822, Sept. 21, 1929. Heavy residues from the distn. of petroleum above 350° or from the thermal dissocn. of hydrocarbons, asphaltic compds., bitumens, tars, tar residues, natural waxes, etc., are converted into a mixt. of gasoline, kerosene and gas oil distg. below 350°. The products to be treated are submitted, under atm. pressure, to the action of an alk. earth base between 350° and 400° so that the gas oil and kerosene products are formed at a temp. below the critical temp. of dissocn. The process may advantageously be combined with a cracking under pressure. The residual oil is transformed either directly or indirectly into gasoline, gas oil and kerosene and the latter 2 are sent either directly or indirectly into the cracking circuit under pressures to be transformed into gasoline and residual oil which is treated again.

Non-detonating motor fuels. ARTHUR B. BROWN and FREDERICK W. SULLIVAN, JR. (to Standard Oil Co. of Ind.). U. S. 1,797,819, March 24. A high b. p. hydrocarbon oil such as gas oil is subjected to conversion temps. under pressure in the presence of NH₃ or a nitrogenous org. compd. of the substituted-NH₂ type such as urea, (NH₄)₂CO₃, guanidine or dicyanodiamide to produce cyclic N compds. in the oil, and a fuel of the gasoline b. p. range is distd.

"Antiknocking" motor fuels. I. G. FARBERND. A-G. Brit. 339,637, July 10, 1929. Antiknocking agents for addn. to motor fuels comprise an amine which is alkylated or arylated on its N atom, and at least 5% of Fe carbonyl. Examples are given of the use of dimethylaniline, monoethylaniline and diethylaniline and the use of other amino derivs. and various diluents, etc., is also mentioned.

Motor fuels and lubricants. ALBERT J. DUCAMP. Ger. 520,010, Mar. 26, 1929. Hg(CN)₂ is added to motor fuels or lubricants, to prevent premature combustion and knocking. The Hg(CN)₂ may be dissolved in glycerol and the soln. dispersed in a hy-

hydrocarbon fuel with the aid of an oil and alkali, or the salt may be dissolved in FeOH and Fe_2O_3 and the soln added to a hydrocarbon fuel, or the salt may be dissolved in an ester, e.g., AcOEt , and the soln added to an alc fuel. An addn of 0.1–1 g of $\text{Hg}(\text{CN})_2$ to 10 l of fuel is suitable. Alternatively, the salt may be dissolved in glycerol, with addn of PhOH if desired, and the soln dispersed in the lubricating oil for the motor.

Lubricants. LUBRICATING PRODUCTS S A and J WOUTHUYSEN Belg 372,620, Sept 30 1930. Lubricants consist of an emulsion of fatty acid, glycerol, NaOH , Na_2CO_3 , AcOH , $\text{K}_2\text{Cr}_2\text{O}_7$, talc, graphite, mineral oil, vegetable or animal oil and water.

Cooling system for lubricating oil such as that of internal-combustion engines. T R CAYE BROWN CALIF Brit 339,650, Sept 10, 1929. A cooling coil mounted in an oil tank contains a liquid of lower b.p. than that of the oil, such as CCl_4 or EtBr , and this coil is connected to a condenser. Various structural details and modifications are described.

Regenerating lubricants such as those from marine or internal-combustion engines. SWAN HUNTER & WIGHAM RICHARDSON, LTD., and HORACE J YOUNG Ger 517,848, Mar 25 1928. See Brit 292,300 (C A 23, 1502).

Filter for waste lubricating oils, etc. STANLEY PIPEKIN (49% to James Higgins) U S 1,798,031, March 24. Structural features.

Transformer and switch oils. I G FARBERNIND A G (Richard Michel, inventor) Ger 519,069, July 17, 1927. Liquid polyalkylnaphthalenes b.p. above 160° are used. The oils may be prepd by condensing C_{10}H_8 or its non-oxygenated derivs with olefins at raised temp. and pressure in the presence of AlCl_3 , and fractionally distg the products.

Asphaltic compositions. HERMANN H. SCHRÖDER Ger. 519,031, Oct. 29, 1926. See Brit 260,621 (C A 21, 3443).

Bituminous emulsions. JONATHAN PARKER Fr. 606,277, May 28, 1930. See Brit 333,303 (C A 25, 589).

23—CELLULOSE AND PAPER

CARLETON E. CURRAN

An explanation of the swelling effect of thiocyanates on cellulose; compounds of cellulose with neutral salts. J R KATZ AND J C DERSEN. *Rec trav chim* 50, 149 52(1931).—The swelling of stretched ramie fibers by coned LiSCN solns has been investigated by the x ray method. When solns more coned than $3\text{LiSCN} \cdot 1\text{H}_2\text{O}$ are used the fibers show an altogether new Röntgen spectrum for cellulose. With solns of lower concns, the salt appears to be adsorbed on the surface of the micelle, without any change in the structure of the latter. With higher concns, presumably addn products are formed on the O atoms of the cellulose, and these compds are responsible for the new x ray patterns. Cellulose can be mercerized by solns of salts of high soly, the degree of mercerization increasing with the salt concn and the length of heating.

Cellulose acetate—its production and uses. HERBERT LEVINSTEIN. *The Times Trade and Eng Suppl* 28, No 664, 8–9(1931).—A review. J H REEDY

Nitrocellulose—a basis for plastics. F SPROXTON. *The Times Trade and Eng Suppl* 28, No 664, 10(1931).—A review. E M SYMMES

pH control in rayon manufacture. CHAS. E MULLIN. *Rayon* 10, No 12, 8, 16, 38–9, 44(1930).—The probable advantages of pH control in the processes of rayon manuf are discussed and the glass-cell electrode is suggested for certain detns. C E M

Paper coloring on the calender and in the size tub. P. W. CARR. *Dyestuffs* 31, 38–41(1930).

Making beater-dyed paper fast to light. IVAR EKHOLM. *Dyestuffs* 32, 10–5 (1931).—Suitable dyestuffs are suggested. CHAS. E MULLIN

Electrical materials commonly used—paper. WM C. HIRSCH. *Elec. Mfg* 7, No 4, 55–7(1931).—A brief general outline of paper manuf is given, attention being specially drawn to the prepn and properties of papers particularly useful to the elec industry. The numerous applications of paper in the manuf of motors, condensers, transformers, etc., and as insulation material, as well as for decorative effects, are reviewed.

EDWARD B. SANIGAR

AcOH and cellulose acetate in the United States (PARTRIDGE) 10. Thermoelastic effect in cellulose ester films (McNALLY, SHEPPARD) 2. Structure of celluloid and nitrocellulose and the gelatinizing medium of nitrocellulose as a swelling medium (KATZ, et al.)

2. The importance of the isoelectric point in the manufacture of casein for paper coating (Bell) 18. Automatically controlling density [of wood pulp aqueous mixtures] (U. S. reissue 18,005) 13. Roller apparatus for applying liquid to paper on one side (Ger. pat. 519,404) 1. Rubber latex [in impregnation of paper] (Ger. pat. 519,483) 30. Preserving [paper] (Ger. pat. 514,981) 20. Synthetic rubber [products for manufacture of coatings, films, artificial silk, etc.] (Brit. pat. 310,008) 30.

FAUST, O. Kunstseide. 4th and 5th ed., revised and enlarged. Dresden: T. Steinkopff. 280 pp. About M 13.50, bound, about M 15.

STARK, C. Die Kollodiumwolle. Ihre Herstellung und Verwendung für Zelluloid, Kunstleder, Nitroseide, Nitrolacke, Filme und plastische Massen. Anhang: Andere Zellulosen. Berlin: M. Krayn.

Cellulose. BERNARDINO OGNETTI. Fr. 697,228, June 11, 1930. Cellulose is extd. from vegetable materials, particularly waste wood, by submitting the material to a treatment with aq. NaOH or KOH of 3-10% concn. and in amt. of 3-6 parts by wt. to 1 part of material for 5-12 hrs. at 100-160°. The material is afterward treated with strong mineral acids or by the wash waters from a previous chlorination and then chlorinated with Cl₂ gas.

Digester for cellulose manufacture. EINAR MORTERUD. Ger. 519,110, Mar. 5, 1930.

Regenerated cellulose films. ERNEST MÜLLER. Ger. 520,105, June 4, 1926. The films are given a final wash dried to remove about 80% of their water content and then treated with a softening agent, e. g., glycerol, glucose or maltose. They are then dried completely. The softening agent may be applied in concd. form to one or both sides of the films, or the films may be passed through an aq. soln. of the softening agent.

Press for alkali cellulose. MASCHINENFABRIK M. HAUSER. Ger. 521,362, Nov. 5, 1929.

Cellulose derivatives. LÉON LILIENTHAL. Ger. 519,138, May 30, 1926. See Brit. 252,634 (C. A. 21, 2381).

Cellulose derivatives. LÉON LILIENTHAL. Ger. 521,452, Mar. 28, 1922. Alkali-sol alkyl derivs. of cellulose are prepd. by treating alkali sol. conversion products or derivs. of cellulose with an alkylating agent in such an amt. that not more than 1 mol. of alkylating agent, preferably 0.53-0.66 mol., is present for each mol. of C₁₂H₂₂O₁₁. Suitable cellulose initial materials are crude or purified viscose and cellulose hydrates. The reaction may be effected in aq. alk. soln. Examples are given. Cf. C. A. 24, 3646.

Cellulose derivatives. U. S. INDUSTRIAL ALCOHOL CO. Fr. 695,401, May 10, 1930. In the manuf. of org. derivs. of cellulose such as the acetate, propionate, benzoate or ethers, a volatile liquid is used as a constituent of the reaction mass and this liquid, by the latent heat of evapn., is used to control the temp. of the reaction. A suitable app. is described.

Moldable masses from solutions of cellulose derivatives. I. G. FARBERND. A.-G. (Erich Richter, inventor). Ger. 514,640, July 8, 1925. The cellulose is partly pptd. from the soln. by precipitants which evap. more quickly than the solvent. Thus, a soln. of nitrocellulose and tricresyl phosphate in AcOEt, AcOEt, toluene and benzene is pptd. in a gelatin network by addn. of a soln. of ester gum in AcOEt and alc. The whole is then kneaded to desired shape. Other examples are given.

Cellulose ethers. IMPERIAL CHEMICAL INDUSTRIES, LTD. Fr. 696,158, May 27, 1930. Cellulose is mercerized in 18-20% NaOH soln., pressed and ripened by storing at about 20-30° and for 12-90 hrs. according to the temp. used and the product desired. The product is treated with at least 50 parts of NaOH for 100 parts of cellulose and then by an alkyl chloride such as EtCl or MeCl. The ethers are sol. in ethylene dichloride, or a mixt. of alc. and C₁₂H₂₂O₁₁, or both.

Cellulose ethers. ALEXANDER WACKER GES. FÜR ELEKTROCHEM. IND. G. M. B. H. (Alfons Ascherl and Wolfgang Gruber, inventors). Ger. 520,384, Feb. 16, 1930. About 1% of a carbonate, calcd. on the alkali cellulose, is included in the reaction mixt., and the reaction is interrupted as soon as CO₂ is evolved. This method of detg. the end point of the reaction is particularly useful in the reaction of alkali cellulose with alkyl halides under pressure. An example is given.

Cellulose esters. I. G. FARBERND. A.-G. Fr. 695,283, May 7, 1930. Esters of cellulose or other carbohydrates are prepd. by first partially substituting by an alkoxy group the OH group of the carbohydrate and afterward esterifying the alkoxy compd. obtained. Thus, cellulose is treated with ethylene oxide and afterward esterified to

give a mixed acetic-butyric ester of a partially hydroxyethylated cellulose which is directly sol. in acetone. Other examples are given.

Cellulose esters. I G FARBENIND A G. Fr. 697,156, June 7, 1930. Esterification of esters is carried out in the absence of free fatty acid and in the presence of liquid chlorinated hydrocarbons not miscible with water and b. below 100° such as $\text{C}_2\text{H}_5\text{Cl}$, $\text{C}_3\text{H}_7\text{Cl}$, or $\text{C}_4\text{H}_9\text{Cl}$, acting as solvents. The reaction is carried out at normal temp. and with less catalyst than usual. In a variation of the process the esterification is effected with the anhydride of one acid in the presence of the chlorinated hydrocarbon and another fatty acid.

Concentrated colloidal dispersions of cellulose esters. HUGH M. SPENCER. U. S. 1797,843, March 24. A cellulose ester such as a cellulose nitrate is incorporated with a substantially anhyd. alkali in soln. such as NH_3 in EtOAc .

Cellulose acetate. KURT BRÄTHER and ALBERT GIERMANN. Ger. 519,137, Feb. 19, 1929. The manuf. of cellulose acetate is effected more rapidly and uniformly by transferring the reaction mixt. from a container under pressure to an evacuated container through a narrow orifice. The mixt. may be repeatedly transferred.

Precipitating cellulose acetate. VEREIN FÜR CHEM. IND. A-G. Ger. 521,125, Aug. 12, 1928. Cellulose acetate soln. and a precipitant are drawn through sep. conduits by a centrifugal pump, which mixes the soln. and the precipitant and feeds the mixt. to a collecting vessel.

Films of cellulose acetate. SRIKES, LTD. Fr. 696,003, June 3, 1930. A plasticizer for use in the manuf. of plastic materials having a basis of a cellulose ester is composed of an org. deriv. of phosphoric acid contg. a naphthyl group, e. g., β naphthyl-diphenyl (or dicycyl) phosphate.

Disintegrating vegetable fibers. FIRMA OTTO C. STRECKER. Fr. 697,030, May 30, 1930. In the disintegration of vegetable fibers and transformation of phenols by the method of phenolates to obtain cellulose, the aerobic enzymes and like organisms of the vegetable fiber used are rendered inactive by heat. The lye contains with the phenolate, free phenol and free phenolic compds. During the cooking the greater content of phenolate is reached before the highest temp. is reached. The gases and hydrocarbons liberated during the cooking are eliminated by distn. from the lye.

Carbohydrate xanthates. WM. HARRISON. Ger. 519,050, Oct. 5, 1928. See Brit. 264,261 (C. A. 22, 164).

Sodium hydrosulfide solutions for viscose manufacture. HOBBS & CO. Fr. 696,153, May 27, 1930. Solns. of NaOH produced by electrolysis with a Hg cathode for the manuf. of viscose are improved by adding to these solns. a small quantity of compds. of Na, K, Fe or Mn, e. g., NaCl , KCl , KOH , complex compds. of K, ferrites, ferrates and their complex compds.

Viscose. FRIEDRICH O. C. KLEIN. Fr. 696,411, Feb. 14, 1930. Viscose is made by passing cellulose or transformation products of cellulose in a const. uninterrupted course, avoiding all treatment in charges, through app. in which the material is simultaneously treated (e. g., impregnated, disintegrated, ripened, pressed, sulfured, dissolved, filtered, ripened again, deaerated, etc.) and transported. A suitable app. is described.

Hollow fibers of viscose, etc. COURTAULDS, LTD. Ger. 516,572, April 26, 1927. See Brit. 273,506 (C. A. 22, 2056).

Mat fibers, etc., from viscose. SIECKBOEN KUNSTSEIDEN, A-G. Ger. 516,573, Nov. 27, 1928. An emulsion of org. material sol. in viscose, etc., is added to the viscose, etc., so that the O or other chemicals in the viscose, etc., ultimately cause pptn. of the material to give the mat finish. Anthrahydroquinone and the phthalic acid ester of isomylal are mentioned in examples as the org. material added.

Filaments, films, etc., from viscose. I G FARBENIND, A-G (Adolf Kämpf, inventor). Ger. 519,349, Jan. 6, 1929. The pptg. bath contains up to 42% of a mineral acid, called as H_2SO_4 , and is substantially satd. with an org. substance of high mol. wt. CH_3O and naphthalenesulfonic acid. The baths are intended particularly for unripened viscose.

Artificial silk from viscose. VEREINIGTE GLANZSTOFF-FABRIKEN A-G. Ger. 519,233 Jan. 13, 1929. Filaments of less than 6 deniers are prepd. by spinning crude viscose through orifices of the normal size into pptg. baths contg. higher concns. of acid than is normal. Higher concns. of acid are used for finer filaments, and the min. concn. of acid is greater or less according to whether the diameter of orifices is greater or less than 0.1 mm.

Artificial silk, etc., from viscose. I G FARBENIND A-G. Brit. 339,740, Nov. 2, 1928. The process described in Brit. 307,829 (C. A. 24, 240) is modified by the use of an

unripened viscose prep'd from a ripened alkali cellulose. The viscose may contain 10% or more of cellulose and the use of high-cellulose viscoses reduces the diln. of the spinning bath. Filaments of 1 denier may be made by this process, having a dry strength of 3 g. or more per denier and a wet strength of 2 g. or more.

Artificial silk. J. A. SINGMASTER. Brit. 339,603, May 8, 1920. For modifying the luster and covering power of artificial silk (such as that made from nitrocellulose or cellulose acetate) by incorporating in the filaments finely divided inert white pigments, the pigments are first distributed in the solvent used for the soln. of the material and after the addn. of the dispersed pigment the resulting soln. is subjected to filtration, which removes the larger particles and renders the material more homogeneous. Various examples of pigments used, etc., are given. Cf. C. A. 24, 3902.

Artificial silk, etc. I. G. FARBENIND A.-G. Brit. 310,138, Dec. 29, 1928. Artificial silk, artificial horse hair or ribbons are made by pptn. of cellulose hydrate by a wet-spinning process, and the freshly pptd. material is stretched by passing it over one or more rubbing surfaces which are heated, so that damage to the material is minimized. For this purpose there may be used glass tubes or rollers which are internally heated by electricity, gas, steam or liquid, and if desired the surface of the stretching device may be irrigated with water or other suitable liquid at about the same temp. in order to prevent accumulation of crystal products, etc., on the heated surface. The strength of viscose silk is thus improved. Some details of app. are described.

Artificial silk. MARTIN HÖLKEN. Ger. 517,747, Oct. 11, 1928. In degasifying cuprammonium silk spinning soln. by reducing the pressure, the gases are passed over solid superphosphate, etc., so that the NH_3 present is absorbed, forming a powerful fertilizer.

Artificial silk. EUGEN MOSSGRABER. Ger. 517,932, Dec. 14, 1929. A device is described for supplying the viscose, acetylcellulose, etc., to the spinning heads for making artificial silk.

Artificial silk. CRIATILLO (SOCIETÀ ANONIMA ITALIANA PER LA SETA ARTIFICIALE). It. 697,261, June 12, 1930. The brilliancy of artificial silk is diminished by producing a current of air or gas perpendicular or oblique to the threads leaving the nozzle by means of a small propeller blade inside the spinning vessel and in the neighborhood of the nozzle.

Artificial silk. HANS SUTER. Fr. 696,306, May 30, 1930. Filaments obtained by dry- or wet-spinning cellulose derivs., particularly cellulose acetate, are drawn under adequate tension in the presence of swelling agents and org. substances which have the power of diminishing the solvent action of the agents. As swelling agents, $\text{C}_2\text{H}_5\text{Cl}$, CHCl_3 and $\text{C}_2\text{H}_5\text{Cl}$ may be used, and as substances to reduce the solvent action, CCl_4 , C_6H_6 , C_8H_{10} , PhMe , xylene and their homologs may be used.

Artificial silk. SOC. POUR LA FABRICATION DE LA SOIE "RHODIASSETA". Fr. 695,490, May 13, 1930. Mat effects are obtained on cellulose acetate silk by dissolving in the spinning soln. a small quantity of a sol. acid such as stearic acid or a sol. anhydride, dry-spinning the solns. obtained and treating the filaments obtained with aq. solns. of salts or bases which form insol. or difficultly sol. compds. with the acid used. Fr. 695,491 describes the production of mat effects on cellulose acetate threads by dissolving in the spinning soln. non-volatile or high-boiling substances, dry-spinning the solns. obtained and removing the added substances from the finished thread by a washing operation. Salts of alk. earth metals and org. substances such as glycerol or stearic acid may be used.

Artificial silk, etc. HENRY DREVYUS. Fr. 696,219, May 23, 1930. In the manuf. of artificial filaments the spinning soln. in its passage to the extruding nozzles is passed through one or more parts in metal or other impermeable substance pierced by a series of small holes which are not sensibly greater, preferably much smaller, than the extruding orifices. Details of construction are given.

Artificial silk. HENRY DREVYUS. Fr. 697,216, June 11, 1930. Filaments or threads are made by extruding a spinning soln. of a cellulose deriv. through a nozzle and receiving the filaments in an evaporative medium in which an exterior skin like layer is formed, and afterward passing them into a solvent vapor to soften the layer and drawing the filaments thus softened. The filaments are then passed through a second evaporative medium. An app. is described.

Artificial fibers. JOHANN BILLWILLER and JOHANN BILLWILLER. Ger. 517,933, Sept. 3, 1922. Carbohydrates contg. little or no fibrous material are worked up by treatment with disintegrating liquids such as solns. of NH_4OH , NaOH , alkali carbonates, hydrosulfides, H_2SO_4 , H_2S , etc., at temps. above 175° and under pressure, with strong and continuous rotary washing. The product is then spun in the manner of artificial silk. In the example bran is treated with NH_4OH at 280° .

Artificial fibers. ALSA SOC. ANON. Ger 521,055, Oct. 30, 1927. The gas content of hollow artificial fibers is raised by soaking the fibers in a liquid in which gas is then evolved by phys. or chem. means. Thus the fibers may be (1) immersed in cold water and with gas, and the water then warmed or (2) immersed successively in liquids which react to form a gas or (3) immersed in a soln. of an unstable compd. that yields a gas on decomposition e.g., a diazo compd. The method may be used to distend collapsed hollow fibers. Examples are given.

Making films, fibers, etc., from cellulose esters and ethers. I. G. FARBENIND. A. G. (Adolf Kämpf, inventor). Ger 520,411, Aug. 20, 1927. The solvent for the ester or ether contains a substantial proportion of CS_2 . 1 example is given.

Artificial threads, films, etc. OTTO SEIDL. Fr 697,272, June 12, 1930. Solns. of cellulose esters are submitted to the action of a concentrated mineral acid. Thus cellulose acetate in an org. solvent is spun and coagulated in H_2SO_4 contg. more than 30 and less than 80% monohydrate. The H_2SO_4 contains at least 0.01% ClH_3O .

Ribbons from cellulose esters or ethers. HENRY DREYFUS. Fr. 695,843, May 19, 1930. Solns. of cellulose esters or ethers, such as cellulose acetate in org. solvents miscible with water, are passed through narrow and relatively long extruding orifices and the ribbons received in coagulating baths composed of water, aq. solns. of salts or aq. solns. of mineral acids.

Artificial silk filaments. I. G. FARBENIND. A.-G. (Otto Faust, inventor). Ger 517,771, Mar. 5, 1927. Filaments which will dye uniformly are spun by passing the filaments from the nozzle over a spring abutment and winding them on a roller, the tension of the spring ensuring a uniform filament.

Spinning nozzle for artificial silk. HENRI COLOMB. Ger 519,348, Nov. 17, 1928.

Spinning nozzle and holder for artificial silk. OSCAR KRONER & Co and ALFRED LEHNER. Ger 516,571, Jan. 24, 1930.

Spinning nozzle with means for uniform aeration of the threads leaving the nozzle. I. G. FARBENIND. A.-G. Fr 697,260, June 12, 1930.

Spinning pot for artificial fibers. SIEMENS-SCHUCKERTWERKE A.-G. (Hans Dietrich, inventor). Ger 520,172, Jan. 28, 1930.

A spinning pot for artificial silk. EDWARD SVATEK. Fr. 695,840, May 19, 1930. Use is made of artificial resin mixed with detached particles of fibrous material and strengthened with layers of fiber soaked in artificial resin.

Continuous filtration apparatus for solutions for spinning artificial fibers. I. G. FARBENIND. A. G. (Adolf Kämpf, inventor). Ger 520,296, Mar. 21, 1928.

Apparatus for making artificial silk by dry-spinning. ADOLF EWALD GULL. Ger 520,479, Oct. 5, 1928. See Brit. 303,778 (C. A. 23, 4573).

Apparatus for making artificial silk from cuprammonium cellulose solutions by stretch-spinning. GACNERT & GIANNETTI. Ger 520,063, Oct. 9, 1925. Brit. 263,462 (C. A. 22, 165).

Apparatus for stretch-spinning of artificial threads. COURTAULDS, LTD., W. H. GLOVES and G. D. BOND. Brit. 339,745, Nov. 4, 1929. Structural features.

Apparatus for making artificial silk filaments, ribbons or straw, etc., by the dry or evaporative process. BRITISH CELANESE, LTD., A. H. TIDMUS, F. A. REBSON and R. H. J. RILEY. Brit. 339,670, Sept. 5, 1929. Two spinning jets are arranged in series and in proximity to each other on a single supply pipe so that the second jet may be removed or replaced without interfering with the first jet. Various structural details are described.

Apparatus for washing and after-treating artificial fibers in spinning pots. I. G. FARBENIND. A.-G. (Richard Hofstadt and Wilhelm Eller, inventors). Ger 521,010, Sept. 16, 1927.

Thread guide for artificial silk. COMPTON DES TEXTILES ARTIFICIELS (SOC. ANON.). Fr 696,211, May 28, 1930. Construction is described of a slightly conical thread guide and drawer mounted in a box where the drops of coagulating liquid are collected and returned to the bath.

Apparatus for treating cakes of viscose silk with liquids. ACME RAYON CORP. Fr 696,965, May 14, 1930.

Washing apparatus for yarn spools. HARZER MASCHINENWERKE G. M. B. H. Ger 517,733, April 17, 1927.

Grinding mill for wood pulp, cellulose, etc. MAX TAMASCHKE. Ger 520,423, Nov. 18, 1927.

Fibrous pulp from wood and other materials. -O. A. MÖLLER. Brit. 340,164, Feb. 9, 1929. Easily bleached cellulose is made from wood or the like contg. cellulose

and lignin by treating the material, after chlorination while moist, with a mixt of a water-sol org solvent such as alc and an inorg acid such as HCl, by using boiling alc. on the countercurrent principle the chlorolignin is extd and the cellulose can be purified by further treatment with fresh alc. MeOH, acetone, glacial HOAc, PrOH, phenol, etc., also may be used, and various details of procedure are described.

Treating waste liquor from chemical wood-pulp manufacture. LINN BRADLEY and EDWARD P. MCKEEFF (to Bradley McKeeff Corp.) U. S. 1,797,678, March 24. The inorg content with admixed org matter of a coned mixt of the waste residual liquor resulting from the digestion of cellulose material such as wood chips in a soln of Na-S salts (contg most of the Na in the form of a sulfite of Na) and the residual liquor resulting from the digestion of cellulose material in an alk soln contg Na compds (the major portion of which are S-free Na compds such as NaOH) is subjected to furnacing under reducing conditions to produce a furnace product contg the Na compds mainly as Na_2CO_3 and Na sulfide, and the resulting Na compds are treated to produce an alk cooking liquor for further use as a digesting agent for cellulose material. Cf C A 24, 4158.

Treating waste liquors such as those of soda and sulfate pulp manufacture. DAVID D. PERRLER (to Pecliles Processes, Inc.) U. S. 1,797,585, March 24. The liquor is dehydrated to a high concn by continuously circulating the heated liquor through a closed expansion chamber while simultaneously continuously removing evolved vapors from the chamber, and the coned waste liquor is turned in a furnace with gases evolved from the chamber.

Rotating autoclave for the extraction of glucose from wood pulp. SILLER & RODEV-KIRCHEN G. M. B. H. Ger 517,890, June 14, 1928. Structural details.

Paper pulp. MEAD PULP & PAPER CO. Brit. 339,599, Aug 28, 1929. Material such as wood, straw, fibrous grasses and the like is used for the prepn of a free pulp of high strength, good color and clay carrying properties, by first subjecting the material to an incomplete chem digestion (suitably for 1-2 hrs at 125° and then at $160-180^\circ$ for 1-4 hrs with addn of CO_2 in a soln of alkali metal sulfite and NaHCO_3 or a tartrate, citrate, borate or oxalate) while maintaining substantial neutrality of the cooking liquor, and then subjecting the material to a limited chlorination treatment while dild and subsequently to a stronger chlorination while thickened. The process is especially suitable for treating waste wood from furniture factories. App and various details of procedure are described. Cf C A 24, 4392, 25, 2563.

Paper. DONALD K. PATTILLO (to Mathieson Alkali Works). U. S. 1,797,789, March 24. Cl is added to the stock in the beater near the beginning of the beating operation to sterilize the stock, which is then subjected to beating and made into paper. Cf. C A 24, 4932.

Paper-making apparatus. HERMAN L. KUTTER (to Black-Clawson Co.). U. S. 1,797,290-1, March 24. Various details of docloe construction, etc., are described.

Paper-making apparatus. T. D. NUTTALL and BENTLEY & JACKSON, LTD. Brit. 310,087, Nov 6, 1929. Structural features.

Paper-making apparatus. C. G. HAUBOLD A.-G. Ger. 520,122, Sept. 6, 1927. Means is described for controlling the run of the wire cloth and the felt.

Dewatering machine for paper manufacture. EISENGESSNER, MASCHINEN- UND PAPPENFABRIK I. A. MCENZNER G. M. B. H. Ger 521,172, Feb 15, 1929.

Couch press for paper manufacture. JOHANNES RAITH. Ger 519,111, Jan 20, 1928.

Suction rollers for dehydrating paper, cellulose, wood pulp, etc. JULIUS STEPHANSEN. Ger 517,661, April 27, 1926.

Machine for disintegrating paper, cardboard, etc. LEOPOLD SCHUMANN. Ger. 516,505, Mar. 19, 1929.

Sizing papers and cloth. JULIUS GLATZ. Fr 695,626, May 15, 1930. Paper and cloth are sized by adding, during or after their manuf., the NH_4 , Na or other salt of the glycolic acid of cellulose sol in water, or homologous compds, or a substance such as starch having an affinity for cellulose.

Water-fast colored paper. SOC ANON POUR L'IND CHIM À BÂLE. Swiss 143,416, April 4, 1929. The paper is immersed in a dye bath which contains also a salt of oleyl-diethylethylenediamine and then is dried.

Corrugated cardboard. THE WESTON PAPER AND MANUFACTURING CO. Ger 519,209, Jan 12, 1930. The cardboard is prepd from a mixt of straw cellulose that has been boiled with Ca(OH)_2 soln and straw cellulose that has been boiled with NaOH soln.

24—EXPLOSIVES AND EXPLOSIONS

CHARLES E. MCKEY AND C. G. STORM

Consumption of explosives in February, 1931. W. W. ADAMS AND L. S. GERRY. Bur. Mines, *Rept. of Investigations* 3163, 15 pp (1931)

ALDEN H. EMERY

The velocity of the phenomena involved in the detonation of solid explosives. P. LAFFITTE AND M. PATEY. *Compt. rend.* 192, 744-6 (1931)—An optical method (C. A. 25, 1580) was used for measuring the velocities of the shock wave and of the luminous gases from the detonation of 10-cm. columns of dynamite A-1 and Astryl, 15 mm. in diam., loaded in tubes at different densities (Δ). The 2 velocities are identical in the vicinity of the charge and are greater than the velocity of detonation in the column of explosive itself. The velocity of the gas then diminishes more rapidly than that of the shock wave, the septa of the 2 waves occurring at the following distances from the charge for tetrel, with $\Delta = 0.45$, 55 cm., $\Delta = 0.65$, 45 cm., $\Delta = 0.91$, 55 cm., for dynamite, with $\Delta = 0.91$, 55 cm., $\Delta = 1.25$, 50 cm., $\Delta = 1.65$, 58 cm. These distances are independent of the length of the column of explosives, but increase as the diam. of the column increases. C. G. STORM

The propulsive strength and rate of pressure development of the Cardox blasting device. N. A. TOLCH AND G. ST. J. FERRITT. Bur. Mines, *Rept. of Investigations* 3064, 7 pp (1931)—The propulsive strength of the Cardox blasting device Model G does not increase directly with the CO₂ charge, but depends on the heater and disk used. In general, the propulsive strength increases with the wt. of the heater and always with the thickness of the disk. The smaller grain sizes of NaClO₄ as used in the heater element gave greater propulsive strength than the larger sizes. An increase in the charge of CO₂ lengthened the duration of the pressure period. The shearing of the disk took place before the pressure was completely developed. With 104-g. heaters and a non-shearable disk the av. max. pressures developed for 570-, 680- and 1155-g. charges of CO₂ were 817, 1280 and 2050 kg./sq. cm., resp.; with 140-g. heaters, 340- and 680-g. charges of CO₂ gave av. max. pressures of 940 and 2110 kg./sq. cm. ALDEN H. EMERY

Ignition, combustion and explosion. N. ARYANOFF. *Chem. Reviews* 11, 103-110 (1931). See 15, 57-8 (1931)—An address. S. S. DE FELIX

Instability of colloidal powders. HENRI MITRAORE. *Sell. soc. chim.* 49, 276-9 (1931), cf. C. A. 25, 817—A further discussion relative to the effect of water and acid produced by the hydrolysis of nitrocellulose on thermolysis is accompanied by comments of de Brum, which are quoted. At temps. of 15-20°, water exerts no action on nitrocellulose, even during periods of 24-30 yrs.; at 50-75°, however, hydrolysis proceeds rapidly in the presence of moisture. Nitrocellulose (13.24% N) kept in 20% HNO₃ for 6 months was reduced to 12.91% N. Thirty % HNO₃ caused the same effect in 3.5 months. As a result of the greater effect of acid of increased concn., a decrease in the moisture content of a powder may, even a trace of free acid may cause an increase in the rate of decompn. C. G. STORM

Influence of moisture on the speed of combustion of colloidal powders. HENRI MITRAORE. *Sell. soc. chim.* 49, 278-84 (1931)—The influence of moisture in smokeless powder was calculated from the ignition temp. (approx. 200°) and the sp. heat of the powder (0.27 for poudre B, and 0.357 for the non-solvent type of powder). The addition of 1% moisture caused an increase of 13.3 and 10.7%, resp., in the energy required for combustion of the two types. Spd_{100} , or the total area of the time-pressure curve, is a measure of the total energy required for decompn. of the powder. Spd_{100} was therefore determined in the Vieille bomb on powder samples with different moisture contents, and the effect of 1% moisture was noted. These expts. showed an increase in Spd_{100} for 1% moisture equal to 12.0% for poudre B, and 10.8% for the non-solvent type of powder. The sample of the latter type contained nitrocellulose (11.7% N), nitroglycerin (approx. 25%) and diethyldiphenylurea, with or without Et phthalate. C. G. STORM

Burning characteristics of smokeless powder. I. Burning temperature. A. M. BALL. *Ind. Eng. Chem.* 23, 498-501 (1931)—The method of calcn. of burning temp. from the heat of combustion of the powder and sp. heats of the gases is discussed in detail. Heats of combustion of the various constituents of smokeless powders are listed, and a tabulation given showing for 16 different powder compos. their heat of combustion, burning temp., gas vol. and compn. of gases, illustrating the effects of increase in content of nitroglycerin and of addition of deterrent materials. C. G. S.

The question of gas formation in city sewers (KEER, MILLER) 14.

GECK, WALTER H.: Die Verhütung von Stauber Explosion. Ein Merkbuch für jeden Betriebsleiter. Berlin: J. Springer 67 pp. M. 6 00

Explosives. HANS RATHENBURG Ger. 521,034, Mar. 17, 1928. See Brit. 308,170 (C. A. 24, 242).

Explosives. DYNAMIT-ACTIEN-GES. VORM ALFRED NOBEL & Co Fr. 690,632, June 4, 1930 The nitrate of ethanolamine and the trinitrate of diethanolamine or nitrates of the homologs of these compds or the N-alkyl derivs of the mono- or di-ethanolamine are used, either alone or with other compds, as explosives. Cf. C. A. 24, 4397

Explosives. WILLEM P. JORISSEN and JOHAN BOOY. Fr. 695,529, May 13, 1930 The explosiveness of detonating mixts., whether gaseous, liquid or solid or as a dispersion, is reduced by adding to the mixts. a gaseous, liquid or solid "anti detonator" such as one or more org. halogen derivs., or one or more phenolic substance or substances which are both halogen and phenolic derivs.

Explosives. LOUIS V. RADUT. Fr. 690,911, Jan. 17, 1930 An explosive suitable for mines and quarries contains NaClO_4 77.1, dinitrotoluene 17.1, castor oil 5.00 and paraffin 0.75%. Fr. 690,912 describes a similar explosive contg. KClO_4 75, dinitrotoluene 18.4, mononitronaphthalene 1.00, castor oil 4.8 and paraffin 0.74%.

Explosives. ALFRED STETTBACHER Fr. 697,213, June 11, 1930 See Brit. 312,310 (C. A. 24, 907)

Blasting fuses and detonators. I. G. FARBENFABRIK A.-G. Brit. 340,041, Sept. 24, 1929 A mixt. of chloronaphthalene (contg. Cl 45-60%) with up to about 20% of S is used for sealing elec. blasting fuses and detonators

Detonators. WILHELM F. SCHUBACH and WALTER FRIEDERICH. Fr. 690,663, June 5, 1930 The capsules for detonators of PbN_3 or $\text{C}_{12}\text{H}_5\text{N}_3\text{O}_4$ are made of an alloy contg. Cu up to 90-95 and Al 5-10%. The latter may be replaced in part by Mg

25—DYES AND TEXTILE CHEMISTRY

L. A. OLNEY

The numerical solubility of dyes. W. C. DUFFER. Proc. Am. Assoc. Textile Chem. Colorists 1931, 103-4; Am. Dyestuff Repr. 20, 185-6 (1931). R. H.

The dyestuffs act in Britain. E. F. ARMSTRONG. Ind. Eng. Chem. 23, 575-80 (1931). G. G.

Fast-to-light substantive dyes. T. F. MURPHY. Dyestuffs 31, 97-9 (1930).—The light-fastness properties of the Solantine dyes are discussed. CHAS. E. MULLIN

Azo-chromophore. III. J. S. P. BLUMBERGER. Chem. Weekblad 28, 190-8 (1931); cf. C. A. 22, 3824—Absorption spectra of about 20 dyes of the α - and β -amino- and -hydroxy-azobenzene types were detd. and compared. The effect of the auxochrome is influenced by the polar character of the substituents, through an inductive action over atoms of alternating functions. In the abnormally behaved α -hydroxyazo dyes the effect on the chromophore is heightened by a coordinate bond between the OH and N N groups. In a series of related, isomeric dyes, the electrostatically most unsym. mol. has the deepest color. The influence of the substituents, whether NH_2 , OH, NO_2 or Cl_2H , depends on their position. A β - NO_2 is strongly bathochrome, a m - NO_2 feebly so in aminoazobenzenes, and hypochrome in hydroxyazobenzenes. A second auxochrome in the primary dye component has a negligible effect; β , β' -diaminoazobenzene and β -hydroxy- β' -aminoazobenzene absorb very much the same as aminoazobenzene, the stronger auxochrome, in the second case, detg. the character of the dye. K. H. ENGEL

Azo dyestuffs derived from arsanilic acids. VENANCIO DEULOFEU and FULGENCIO MARIN. Anales soc. españ. fis. quim. 29, 65-73 (1931).—By diazotizing the acids and coupling with the phenols in alk. soln. the following compds. were prepd.: From β -arsanilic acid and resorcinol, 2,4-dihydroxyazobenzene-4'-arsonic acid, yellow-orange, with phloroglucinol, 2,4,6-trihydroxyazobenzene-4'-arsonic acid, red brown, with orcinol, 2-methyl-4,6-dihydroxyazobenzene-4'-arsonic acid, deep red brown; from 2-methyl-4-aminobenzenearsonic acid and resorcinol, 2,4-dihydroxy-2'-methylazobenzene-4'-arsonic acid, orange; from α -arsanilic acid and PhOH, 4-hydroxyazobenzene-2'-arsonic acid, yellow; with resorcinol, 2,4-dihydroxyazobenzene-2'-arsonic acid, light orange-yellow; with phloroglucinol, 2,4,6-trihydroxyazobenzene-2'-arsonic acid, red; and with

oreinol, 2-methyl-4,6-dihydroxyazobenzene 2' arsonic acid, deep red. Comparison of the colors of the dyes shows that they follow the Nietzki rule. F. M. SYMES

Properties of insoluble azo colors on the fiber in relation to their constitution. Fastness of kier boiling and the method of testing. F. M. ROWE. *J. Soc. Dyers Colourists* 47, 29-31(1931).—The products of the interaction of β -hydroxynaphthoic acid and aromatic amines depend upon the conditions employed. Examples of the reaction are given. Halogen and other groups in the compounds obtained from this reaction increase the fastness. The presence of a nitro group in an azo coloring matter has an adverse effect on the fastness or kier boiling of the dyeing. ROBERT HOUGHTON

Action of boiling caustic soda (kier boiling) on insoluble azo colors on the fiber. F. M. ROWE. *J. Soc. Dyers Colourists* 47, 35-9(1931).—Cotton exerts a favoring influence to decomposition of the color aggregates into other compounds. A study is made of the effect of boiling NaOH upon the chemical decomposition of the color lakes of some azo colors. ROBERT HOUGHTON

New kinks in wool dyeing. WERNER LOV BERGEN, *Proc. Am. Assoc. Textile Chem. Colorists* 1931, 94-103. *Am. Dyestuff Repts.* 20, 180-5(1931).—The effect of metals on wool dyes causing change of color and a method for detecting this effect are presented. Vat colors are dyed on wool at pH 9.2-9.1. Tippy dyeings and methods of avoiding them are discussed. Chlorinated wool can be detected by treating under the microscope with 0.1 N NaOH. A decided swelling indicates chlorinated wool. ROBERT HOUGHTON

Dyeing of union felt. WALTER WINKLER. *Dyestuffs* 31, 41-2(1930). C. E. M.

The dyeing of rayon yarns. D. O. GRIDER, *Proc. Am. Assoc. Textile Chem. Colorists* 1931, 92-8, *Am. Dyestuff Repts.* 20, 174-80(1931).—Methods of handling rayon yarns during dyeing are discussed. ROBERT HOUGHTON

Dyeing raw stock for carpet yarns. F. C. WALSH. *Dyestuffs* 31, 186-7(1930). CHAS. E. MULLIN

Dyeing carpet yarn. HARRY PEERS. *Dyestuffs* 31, 28-9(1930).—Formulas and dyes are suggested. CHAS. E. MULLIN

Dyeing cotton yarn in the package. KENNETH MACKENZIE. *Dyestuffs* 31, 12-4(1930).—The application of the direct, developed, S and vat dyes is discussed. CHAS. E. MULLIN

Printing of cotton. GEO. M. BISHOP. *Dyestuffs* 32, 1-10(1931).—Direct, discharge, vat, and naphthol printing are discussed, with formulas. CHAS. E. MULLIN

Wooden vessels in the dyehouse. FRED GROVE PALMER, *Am. Dyestuff Repts.* 20, 189-93(1931).—The selection of the proper wood for the use and installation are discussed. ROBERT HOUGHTON

Peculiarities of fading. WM. H. CADY, *Proc. Am. Assoc. Textile Chem. Colorists* 1931, 120-5, *Am. Dyestuff Repts.* 20, 232-7(1931), cf. *C. A.* 25, 386.—Combinations of fast and fugitive dyes are given which exhibit peculiar fading properties. Various methods of fading by light exposure are presented and discussed. R. H.

Detection of traces of active chlorine in textiles. ALFRED SCHMIDT. *Z. angew. Chem.* 44, 278(1931).—The iodo-starch test recommended by K. Schwarze (*C. A.* 25, 1387) is inconclusive, inasmuch as chloramines present in vegetable fibers will give the test. Whether they are injurious, or to what degree, is not yet known with certainty. They can be destroyed by treatment with bisulfite or thiosulfate but are more stable than hypochlorites, which react with hot Na₂O₂ solutions. W. T. H.

Washing of wool and its chemical control. V. DÖHRMANN. *Kisérlet Közlemények* 32, 405-13(1929).—Well washed wool should not contain more than 1% fat. No washing agent should remain in wool and washing should not deteriorate the physical qualities of wool fibers. Na₂CO₃ solution weaker than 2.5% combined with a neutral K soap should be used as washing substance at exactly 45°. Washed wool should be examined for (1) fat content by extraction, (2) content of Na₂CO₃ by titration, and (3) quality of wool fibers by the elasticum reaction of Allwarden. S. S. DE FIVELY

Artificial wools and wool-like finishes on fabrics. CHAS. E. MULLIN AND FLORENCE H. CADWELL. *Textile Colorist* 52, 161-4, 193, 241-4, 269(1930).—A review and discussion of the patents and literature on the manufacture, uses, properties and probable future of the so-called artificial wools, wool substitutes and imitation wool finishes on the vegetable fibers. CHAS. E. MULLIN

Identification of artificial silks. A. K. GYZANDER. *Dyestuffs* 31, 1-2(1930).—National Erie Black GXOO dyes the regenerated silks black and acetate silk yellow or orange. A solution containing 20 mg. National Erie Yellow F or National Erie Fast Orange CG, 50 mg. Na₂CO₃, 1 g. soap and 1 g. Na₂SO₄ in 50 cc. H₂O dyes cuprammonium silk in a few minutes but viscose is hardly stained. CHAS. E. MULLIN

The investigation of faults in rayon hosiery. V. T. F. HEYES. *Rayon Record* 5, 233-7 (1931), cf. C. A. 24, 5504—11 gives methods for the detection and examn of the character of oil and metallic soaps in rayon hosiery. A. K. JOHNSON

Peraktivin. RICHARD FEINELMANN *Chem.-Ztg.* 55, 191-2 (1931).—In order to strengthen the bleaching power of aktivin, dichloramine is suggested in place of chloramine. Its insoly in H₂O is disadvantageous, but it can be used in a finely divided or colloidal state, by dissolving in alkali and diluting. Peraktivin contains about 30% active Cl, and is employed as a bleaching agent in the textile industry and as a "safe" bleaching agent in laundries. CURTIS L. WILSON

Waste gas-cleaning materials [from dyes works] as weed destroyer [ANON.] 15. Stream pollution by textile wastes [WESTON] 14. Treatment of a sewage containing wool-scouring refuse [BERNHAM] 14. Anthracene derivatives [dyes] (RUGGLI, HEITZ) 10. Dyes derived from acenaphthenequinone (GUTH) 10. The structure of aniline black [JOPPE, METRIKINA] 10. Anthraquinone azo compounds 11. Insoluble disazo dyes from 1,5-diaminoanthraquinone (MAKI) 10. The natural coloring matter of raw silk fiber of the domestic cocoon (OKU) 11D. Automatically controlling density or character [of dye solution] (U. S. reissue 18,005) 13. Rubber latex [in impregnation of fabrics] (Ger. pat. 519,483) 30. Agents [for use in dyeing] (Fr. pat. 696,327) 29. Adhesive [for patching fabrics] (Ger. pat. 521,199) 18. Zn-Ca formaldehyde sulfoxylate [for printing textiles] (Ger. pat. 521,201) 28. Organic P compounds [for combating moths] (Ger. pat. 521,205) 17. Sizing cloth (Fr. pat. 695,626) 23. Condensation products containing halogen [starting materials for manufacture of dyes] (Fr. pat. 695,602) 18. Colored condensation products [wool dyes, intermediates for dyes] (Ger. pat. 519,051) 10.

LEHNE, A., AND HEGEL, K. T. *Textilchemische Erfindungen*. Lfg. 7. Wittenberg. A. Ziemsen Verlag. 172 pp. M. 10. Cf. C. A. 25, 1390.

SCHULTZ, GUSTAV. *Farbstofftabellen*. 7th ed. Band I. Lfg. 14-15. Leipzig. Akad. Verlag. G. m. b. H. Pp. 677-672. Cf. C. A. 25, 1390.

Dyes. I. G. FARBENIND. A.-G. (Karl Wilke, Josef Stock and Fritz Schubert, inventors). Ger. 516,698, Dec. 2, 1928. Dyes of the benzanthronepyrazole-anthrone series are prep'd by the action of aminobenzanthronepyrazoleanthrone with compds which cause a substitution of the N-linked H atoms. Numerous examples are given, in which the amino compd is treated with As₂O₃, BzCl, cyanuric chloride, Na₂P₂O₇, Me₂SO₄, 1-chloroanthraquinone and AeONa, and other substances, to produce dyes of various colors.

Dyes. SOC. ANON. POUR L'IND. CHIM. A. BALE. Swiss 143,025, Mar. 16, 1929. A new dye is obtained when 2 mol. 4-aminoazobenzene-3'-sulfonic acid is condensed with 1 mol. of 4,4'-diminotrostilbene-2,2'-disulfonic acid under pressure. The dye colors cotton in light- and alkali-fast yellow-orange shades.

Azo dyes. I. G. FARBENIND. A.-G. Brit. 340,009, Aug. 10, 1929. Azo dyes of the tartrazine series are obtained by a selection of the reacting components such that the finished dye is substituted in the 1- and 4-positions of the pyrazolone ring by a phenyl and a phenylazo radical, resp., which each contain 2 sulfo groups, and such that of the 4 sulfo groups one pair occupies the 5'- and 5"-positions while the other pair occupies either the 2'- and 2"- or the 3'- and 3"-positions, or, if each of the phenyl radicals carries a methyl group in the 2'- and 2"-positions, the second pair of sulfo groups occupies the 4'- and 4"-positions. The products give greenish yellow dyeings fast to water and are especially suitable for printing on paper and are convertible into lakes by the usual methods. Numerous details and examples are given.

Azo dyes. I. G. FARBENIND. A.-G. Brit. 339,620, July 10, 1929. Azo dyes insol. in water are made in substance or on a sub-stratum or fiber such as regenerated cellulose and cellulose-ester and -ether fibers, by coupling a diazo or tetrazo or diazoazo compd free from COOH, SO₃H and OH groups with an arylide such as 2',3'-hydroxynaphthoyl-4-amino-1,2-dimethoxybenzene (which may be made from 2,3-hydroxynaphthoic acid and the corresponding 4-amino-1-alkoxymethoxybenzene). Examples are given of various suitable components.

Azo dyes. I. G. FARBENIND. A.-G. (Hermann Wagner and Karl Beck, inventors). Ger. 520,240, Oct. 22, 1927. The diazo compds from esters of 2,3-aminonaphthoic acid or their substitution products are coupled, in substance or on the fiber, with arylides of 2,3-hydroxynaphthoic acid or their substitution products. Examples are given. Cf. C. A. 25, 2299.

Azo dyes. I G FARBERND A G (Heinrich Clingstein and Karl Wiedemann inventors) Ger 520,242, July 27, 1928 The dyes obtained by coupling diazotized *o*-aminophenolsulfonic acids or their nuclear substitution products with 1,7-aminonaphthol or its *N* acyl derivs are treated in substance with chroming agents, yielding blue dyes Examples are given

Azo dyes. I G FARBERND A-G (Leopold Laska and Arthur Zitscher, inventors) Ger 520,241, May 8, 1928 Diazo, tetrazo, or diazoazo compds not contg the COOH or SO₃H group are coupled, in substance or on a support, with a 2,3 hydroxynaphthoyl 2 amino-5-halohydroquinone dimethyl ether The dyes are insol in water, and are useful as pigment colors as well as for dyeing or printing vegetable fibers Numerous examples are given Cf C A 25, 229

Azo dyes. I G FARBERND A-G (Leopold Laska and Arthur Zitscher, inventors) Ger 520,243, Feb 2, 1928 See U S 1,762,022 (C A 24, 3652).

Azo dyes. I G FARBERND A-G Fr 696,542, May 27, 1930 Dyes contg Cu are prepd by treating dyes obtained by coupling tetrazotized 4,4'-diaminobiphenyl 3,3'-dicarboxylic acids with 2 mols of aminonaphtholsulfonic acids or their substitution products, with agents yielding Cu or by prepg the dyes in the presence of agents yielding Cu The products dye cotton in fast blue and gray shades Several examples are given

Azo dyes. I G FARBERND A-G Fr 696,873, June 10, 1930 Azo dyes, which dye wool in shades going from red to violet, are prepd by introducing acyl radicals into the monoazo dyes obtained by coupling diazotized 4 nitro-1 aminobenzene 2-sulfalkyl, -aryl- or -aralkylamides with 2 naphthylaminosulfonic acids or 2 amino-8-naphtholsulfonic acids or their derivs and reducing in an alk medium (See Fr 515, 412) Examples are given

Azo dyes. Soc ANON POUR L'IND CHIM A BALE Fr 697,141, June 6, 1930 Metallized azo dyestuffs are prepd by treating azo dyes capable of forming complex metal derivs, with substances which yield metals, in the presence of org bases such as pyridine, quinoline, PhNH₂, triethanolamine or benzylamine The azo dyes were derived from coupling components such as naphthylamines of their derivs Examples are given Cf C A 25, 822

Azo dyes. Soc ANON POUR L'IND CHIM A BALE Fr 695,809, May 20, 1930 The azo dyestuffs obtained by the union (in an alk medium) of diazo compds contg a OH or COOH group in the *o*-position with respect to the azo group, with derivs of 2,5 aminonaphthol-7-sulfonic acid substituted at the N atom by an appropriate group such as an alkyl, aralkyl or aryl group, are treated with HNO₃ and agents capable of liberating metals to form lakes The dyes are particularly suited for regenerated cellulose Several examples are given

Vat dyes. C SHAW, R F THOMSON, J THOMAS and SCOTTISH DYES, LTD Brit 339,954, June 11, 1929 *Bz*-1, *Bz*-1'-dibenzanthronyl or a substitution product is treated with a dichromate in H₂SO₄, and the product is subjected to alkali fusion, with or without a diluent such as ale or phenol The resulting dyes give blue-gray to black shades

Vat dyes. I G FARBERND A-G (Erwin Hoffa, Hans Heyna, Erwin Thoma and Otto Hirschel, inventors) Ger 521,051, May 30, 1926 3,5-Dimethyl-4 halophenyl-1 thioglycolic acids, or 3,4-dimethyl-5-halophenyl-1-thioglycolic acids, or their derivs substituted in the 2 position with CN, CONH₂, or COOH group, are converted into the corresponding hydroxythionaphthenes, and these are converted into sym thiondigo dyes by oxidation, or condensed with the customary thiondigo dye components Examples are given

Vat dyes. I G FARBERND A-G. (Werner Zerweck, inventor). Ger. 517,846, Apr 3, 1928 A brown vat dye of the anthraquinone series is produced by condensing aminoanthracene with 1-chloroanthraquinone 2-carboxylic acid and treating the product with an acid condensing agent Thus, naphthalene, 1-chloroanthraquinone 2-carboxylic acid, (COOH), and β -anthramine are heated in a reflux condenser, and the product diluted with ligroin The dye colors vegetable fibers in fast red brown shades Several further examples are given Cf C A 25, 1094

Vat dyes. I G FARBERND A-G (Georg Rösch, inventor). Ger. 520,088, Sept. 13, 1929 Addn to 518 017 (C A 25, 2573) Yellow to orange vat dyes are prepd as in Ger 518,017, using naphthalene-2,7-dicarboxylic acid or its derivs not contg OH groups, instead of naphthalene 2,6 dicarboxylic acid or its derivs Examples are given

Vat dyes. I G FARBERND A-G Fr 696,011, May 23, 1930 Vat dyes of the anthraquinone series are prepd by treating with acid condensing agents the products resulting from the reaction of aminoanthraquinones with 5,8 dihalo-1,2 benzanthra-

quinones. Thus, the product obtained by the reaction of 1 mol of 5,8-dichloro-1,2-benzanthraquinone with 2 mols of 1-aminoanthraquinone is introduced into a soln. of $AlCl_3$ in pyridine and heated to 125–130°.

Vat dyes. I. G. FARBENIND A.-G. Fr. 696,423, April 8, 1930. Vat dyes of the anthraquinone series are prepd by causing one or more radicals of anthraquinone- β -carboxylic acid, substituted or not, to enter into the NH_2 group of α -aminoanthrimid-carbazoles or their derivs or substitution products by known methods. Examples and the formulas of the compds obtained are given.

Vat dyes. I. G. FARBENIND A.-G. Fr. 697,231, June 11, 1930. Acylating agents are caused to react with condensation products, or their derivs halogenated in the benzanthrone ring, which are obtained by the alkali treatment of 1-(benzanthroneyl-amino)-4 (5 or 8)-aminoanthraquinone. Similar dyes are prepd by submitting 1-(β -1-benzanthronylamino)arylaminoanthraquinones or their substitution products to an alk. condensation using agents practically free from OH groups. Several examples are given.

Green vat dyes. I. G. FARBENIND A.-G. Fr. 696,862, June 10, 1930. New dimethoxydibenzanthrone is halogenated in the presence of indifferent org. solvents and catalysts. Examples are given.

Brown vat dye. PAUL NAWIASKY and JULIUS MUELLER (to General Aniline Works) U. S. 1,797,478 March 24. The nitrobenzanthrone obtainable by treating benzanthrone with HNO_3 in boiling glacial HOAc is treated with a caustic alkali. Cf. C. A. 24, 1225.

Dye mixtures. SOC. ANON. POUR L'IND. CHIM. A. BALE. Ger. 521,476, June 19, 1923. Brown dye mixts are prepd by mixing sym. 1,2-naphthothiondigo with sym. 2,1-naphthothiondigo, or by mixing the corresponding leuco compds or hydroxythionaphthenes and oxidizing the mixts.

Metallic dyes. SOC. ANON. POUR L'IND. CHIM. A. BALE. Swiss 143,026, Mar. 16, 1929. The azo dye from 5-nitro-2-diazo-1-phenol and 1-amino-8-hydroxynaphthalene-4-sulfonic acid is treated with Cu yielding agents. The dye colors silk in greenish blue shades.

Metallic dyes. SOC. ANON. POUR L'IND. CHIM. A. BALE. Swiss 143,027, Mar. 16, 1929. The chromed dye from 5-nitro-2-diazo-1-phenol and 2-phenylamino-5-hydroxynaphthalene-7-sulfonic acid is reduced, diazotized and coupled with salicylic acid. The dye colors cotton and artificial silk in blue-green shades.

Wool dyes. I. G. FARBENIND A.-G. Fr. 695,807, May 17, 1930. 1-Amino-4-bromoanthraquinone-2-sulfonic acid is condensed in aq. soln. and in the presence of agents neutralizing the acid and a catalyst with arylidiamino compds. (with the exception of α -diamines and p -phenylenediamine) or their substitution products, particularly their sulfonic acids contg. at least one free NH_2 group. The condensation products obtained may be treated with acylating agents, or the sulfonic groups contd. in the anthraquinone ring may be sepd.

Anthraquinone dyes. IMPERIAL CHEMICAL INDUSTRIES, LTD. Ger. 521,382, July 27, 1929. See Brit. 323,026 (C. A. 24, 2890).

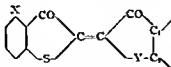
Anthraquinone vat dyes. I. G. FARBENIND A.-G. Brit. 339,626, Aug. 8, 1929. 2,2'-Dimethyl-1,1'-dianthraquinonyls or 2,2'-dimethyl *ms*-benzo- or naphtho-dianthrone in which the H atoms of both methyl groups are partially or wholly replaced by halogen are treated with acid or alk. condensing agents having a reducing action and capable of splitting off halogen or a H halide. Numerous details, examples and modifications of procedure are described. The dyes dye materials orange-yellow to red shades.

Indigoid dyes. COMPAGNIE NATIONALE DE MATIÈRES COLORANTES ET MANUFACTURES DE PRODUITS CHIMIQUES DU NORD RÉUNIES ÉTABLISSEMENTS KUELMANN. Fr. 695,979, May 22, 1930. 4,6-Dimethyl-5-nitro- and 4,6-dimethyl-5-amino-1-phenylthioglycolic acids are obtained by diazotizing 1,3-dimethyl-2-nitro-4-aminobenzene in the presence of a mineral acid, pouring the diazo soln. into a dil. soln. of alkali xanthogenate or the like in alk. medium, saponifying, adding alkali monochloroacetate or the like and acidifying after elimination of alc. The new dyes, 5,7,5',7'-tetramethyl-6,6'-dinitro- and -diamino-2,2'-bis(hianaphthene)indigo, are obtained by cyclization of the acids mentioned, e. g., by means of CSO_2H .

Indigoid dyes. SOC. ANON. POUR L'IND. CHIM. A. BALE. Ger. 519,265, Nov. 16, 1928. New indigoid dyes are prepd from 1-methyl-2,3-naphthoxythiophene by oxidizing this or condensing it with the cyclic diketones or their reactive derivs. usually employed as indigoid dye components. 1-Methyl-2,3-naphthoxythiophene is prepd from 1-methyl-2-aminonaphthalene by diazotization, treatment with alkali xantho-

genate and sapon to yield 1-methyl-2-thionaphthol, conversion of this into 1-methyl-2-naphthylthioglycolic acid, followed by conversion of the acid into its halide and treatment of this with acid condensing agents to effect ring closure. The dyes obtained may be halogenated. Examples are given. *C. C. A.* 25, 825.

Thiondigo dyes. I. G. FARBERND A-G. *Fr.* 695,763, May 16, 1930. 3-Halo-phenyl 1-thioglycolic acids, which are substituted in the 2-position by a carboxylic, carbamate or nitrile group, are transformed into the corresponding hydroxythionaphthenes and the latter by oxidation into symmetrical thiondigo dyes, or the said hydroxythionaphthenes or their 2-derivs may be condensed with diketones or their 2-derivs. Compsd. of the formula



are obtained in which X is a halogen, Y an imino group or an atom of S, and C₁ and C₂ belong to a C₆H₄ or a C₆H₃ ring.

Triarylmethane dyes. I. G. FARBERND A-G. *Brit.* 339,823, Jan. 4, 1929. 1-Naphthaldehyde 2-sulfonic acid or a deriv. is condensed with an alkylated or aralkylated arylamine and the resulting leuco compd. oxidized (if desired after sulfonation). The dye formed from 1-naphthaldehyde 2,4-disulfonate and ethylaniline produces a bluish green color on wool and other examples also are given.

Xanthene dyes. IMPERIAL CHEMICAL INDUSTRIES, LTD. *Ger.* 521,201, June 30, 1929. See *Brit.* 320,345 (*C. A.* 24, 2608).

Xanthene dyes. IMPERIAL CHEMICAL INDUSTRIES, LTD. *Ger.* 519,345, June 30, 1929. See *Brit.* 320,345 (*C. A.* 24, 2608).

Sulfurized phenol derivatives for use in dyeing. I. G. FARBERND A-G and A. THALSS. *Brit.* 340,007, Aug. 19, 1929. In processes such as described in *Brit.* 173,313 (*C. A.* 16, 1510) and *Brit.* 298,280 (*C. A.* 23, 2837) for the production of mordanting thio derivs of phenols by heating phenol or a homolog or substitution product (other than a phenol contg. N) with S in the presence of a neutral salt of an org. carboxylic acid or of a weak inorg. acid, an excess of at least 50% of phenol over the theoretical quantity of 1 mol phenol to 2 mols of S is employed. The reaction proceeds rapidly to completion and the excess of phenol is removed by distn. *in vacuo*.

Bromo substitution product of isodibenzanthrone. SOC. ANON. FOUR L'IND. CHIM. A. BALE. *Swiss* 143,711, July 20, 1929. Isodibenzanthrone is treated with 2 atoms of Br in the presence of ClSO₃H and a carrier such as S or I. The product dyes cotton violet blue from a red blue vat.

Pyrazoleanthrone condensation products. I. G. FARBERND A-G (Georg Kalischer, Heinz Scheyer and Heinrich Rutter, inventors). *Ger.* 517,845, Feb. 17, 1929. 2-Methylpyrazoleanthrone is alkylated or aralkylated, and the N-alkylated or N-aralkylated product treated with alkali until the product gives a red color in concd. H₂SO₄. Thus 2-methylpyrazoleanthrone is treated with Me₂SO₄ to produce 2-methyl-2-methylpyrazoleanthrone, m. 163°. This is heated with NaOH in a reflux condenser to give a dye which colors cotton red from a green blue vat. Further examples are given.

Dye pastes. IMPERIAL CHEMICAL INDUSTRIES, LTD. *Ger.* 521,123, July 6, 1929. See *Brit.* 300,800 (*C. A.* 23, 4079).

Colored lakes fast to light. I. G. FARBERND A-G. *Fr.* 695,495, May 13, 1930. Basic dyes, which may contain acid groups, are treated with complex acids or their salts and substances having a basic reaction in amt. not sufficient for complete neutralization are added before allowing them to react on the dyes. Examples are given.

Dyes and intermediates. I. G. FARBERND A-G. *Brit.* 339,659, July 8, 1929. *ms* Naphthodianthrone, *alle ms* naphthodianthrone and *ms* anthrachanthrone are treated with a nitrating reagent (preferably in the presence of diluents such as mono-, di- or tri-chlorobenzene, bromobenzene or quinoline) and the nitro groups in the products are replaced by halogen or are partly or wholly reduced. Examples are given for producing dyes giving various colors.

Dye intermediates. I. G. FARBERND A-G. *Brit.* 339,699, Sept. 24, 1929. Aromatic aminoaldehyde compds. are made by heating aldehydes of the benzene and anthraquinone series contg. easily replaceable halogen atoms with acid amides or sulfonamides or their monoalkyl, aralkyl or aryl derivs, in the presence of an acid binding

agent, with or without a diluent or catalyst, and saponifying the acylamino or sulfamino compds obtained. 1 samples are given of the production, as final products, of 5-nitro-2 aminobenzaldehyde, 6-methylacridine, 6 chloro-2 amino- and 2,6-diaminobenzaldehyde, 3 nitro-2,6-diaminobenzaldehyde, 2,5-diaminoterephthalaldehyde and 1 aminoanthraquinone 2 aldehyde.

Dye intermediates (anthraquinone derivatives). P. F. BANGHAM, L. J. HOOLEY, J. THOMAS and SCOTTISH DYES, LTD. Brit. 339,589, June 4, 1929. Anthraquinone derivs with heteronuclear substituents are obtained by condensing halogenated phthalic acids, salts or anhydrides with halogenated phenols in H_2SO_4 or oleum with H_2SO_4 or borax. Halogen atoms in a positions may be wholly or partly hydrolyzed to OH groups. The products are chlorinated quinizarins and alizarins. 1 examples are given.

Dye intermediates. I. G. FARBERND A-G. Fr. 696,043, May 23, 1930. Substituted phenol α -carboxylic acids are prepd by heating under pressure with CO_2 the alkali salts of chloromethyl or dimethyl phenols, particularly asym α -xylenol, p -chloro- α -cresol or α -chloro- p -cresol. 1 examples are given.

Dyeing. I. G. FARBERND A-G. (Franz Streng, inventor). Ger. 519,252, July 29, 1928. Colored reserves under aniline black are obtained with a printing paste free from nitro-substituted bases and comprising a nitrosamine alkali salt and, as coupling component, an arylide of 2,3 hydroxynaphthoic acid or a compd contg a reactive CH_3 group. The paste contains also ZnO and an excess of caustic alkali. The material is slop padded with the usual aniline black liquor, dried, printed with a paste of the kind described, steamed to develop the black, and then passed through a bath contg AeOH and a bichromate soda bath. Cf. C. A. 24, 5509.

Dyeing fabrics. KARL SCHMIDT. Ger. 517,910, Oct. 21, 1926. An aniline black dye is produced by a slop pad bath contg PhNH_2 , an oxidizing agent, a large amt of weak acid salt, an org. catalyst and a heavy metal salt as an auxiliary catalyst. Such a bath contains PhNH_2 , HCl , $\text{Na}_2\text{S}_2\text{O}_8$, CuSO_4 , NaClO_4 , $(\text{NH}_4)_2\text{SO}_4$ and p -phenylenediamine. The slop padded goods are dried and developed by a Cr bath. A further example is given.

Dyeing fabrics. SOC. ANON. POUR L'IND. CHIM. A. DALE. Ger. 521,200, Dec. 22, 1928. Color effects on fabrics, etc., are obtained by treating the materials with an alkali salt of a 1 naphthol-4 aryl ketone and with a dye or other compd fixable on the materials by steaming. The materials are then steamed, and finally the ketone is coupled with a diazo compd to produce an azo dye. Thus, a mixed fabric of cotton and acetate silk may be padded with a bath contg a salt of 1-naphthol-4 phenyl ketone and aminoazobenzene, steamed so as to fix the dye on the acetate silk, and then treated with a diazo soln from 4 chloro-2 aminolenzene 1 phenyl ether. The acetate silk is colored yellow and the cotton red. Further effects may be obtained by impregnating the materials with the ketone, printing with dyes requiring steaming, and then proceeding as above.

Dyeing animal fibers. DURAND & HUGUENY S. A. Ger. 516,660, April 21, 1928. Addn to 418,487. Colors fast to rubbing are produced on animal fibers by ester salts of leuco vat dyes. The dyes are developed by acid saponin soaps in an acid oxidizing bath. Thus, wool, soaked in a 10% soln of ester salt of leucoindigo, is immersed in a cold bath contg saponin, NaNO_2 and H_2SO_4 . The wool is washed and dried. Leuco compds of quinone and anthrone are mentioned also.

Dyeing silk. COMPAGNIE NATIONALE DE MATIÈRES COLORANTES ET MANUFACTURES DE PRODUITS CHIMIQUES DU NORD, RÉUNIES ÉTABLISSEMENTS KUHLMANN. Fr. 695,981, May 22, 1930. Natural silk is dyed by combination with diazonium salts of amines such as m -chloroaniline, m -nitro- α -anisidine or bisanisidine.

Dyeing acetate silk. I. G. FARBERND, A-G. Fr. 696,026, May 23, 1930. Acetate silk is dyed by treating it with dyes of the type $(4)\text{H}_2\text{NRN}\cdot\text{NR}'\text{NX}_2(4)'$, diazotizing and developing with 2 hydroxynaphthalene 3-carboxylic acid. In the formula R is a phenyl group substituted or not by halogen, an alkyl or alkoxy group or several of these, R' is a phenyl or naphthyl group substituted or not by a halogen or an alkyl, and X is an aliphatic group. Examples are given.

Dyeing materials containing cellulose acetate and like substances. BRITISH CELANESE, LTD. Brit. 339,997, June 8, 1928. Dispersed relatively insol coloring agents having an affinity for unsaponified cellulose esters are used for dyeing materials such as mixed goods contg partially saponified cellulose acetate or the like which may be assocd with cotton, wool, silk, linen or regenerated cellulose, and the non-ester portion may be dyed with suitable dyes having an affinity for such portion.

Dyeing cotton associated with "cuprammonium yarn." FRITZ STEINBERG (to American Bemberg Corp.). U. S. 1,797,247, March 24. For equalizing the dyeing

effect of a fabric comprising a mixt. of cotton and "cuprammonium yarn," the cotton is treated with a Cu-oxide-NH₃ soln.

Printing on cloth with dyes. I G FARREND A-G (Friedrich Teller, inventor) Ger 510,440, Mar 20, 1929. See Brit. 229,207 (C. A. 24, 5510)

Printing textiles. I G FARREND A-G (Heinz Goeder, inventor) Ger 272,19, Apr. 1, 1928. Reserves colored with vat dyes under dyings with vat dyes are obtained with the use of reserve pastes contg. thickeners which permit the addition of acid reacting heavy metal salts or like reserving substances to the paste, and which fix the reserve dye as a pigment to such a degree that the reserve dye is not removed during the dyeing with the grounding dye. Suitable thickeners are cellulose ethers, carob-bean gum and blood albumin. Thus, the goods may be printed with a reserve paste contg. a rose vat dye, carob-bean gum, CuSO₄, Cu(NO₃)₂, and Cu(OAc)₂, then vat-dyed in the ordinary way with indanthrene blue, washed, acidified, washed again and dried. The reserve color is then fixed by padding the goods with a soln. contg. Rongalite C, NaOH and glycerol, steaming and finishing as usual. Cf. C. A. 24, 974

Use of color reserves in textile printing. I G FARREND A-G (to Durand & Huguenin A-G) Brit. 340,083, Nov. 5, 1928. Color reserves under vat dyings are obtained by means of ester salts of leuco vat dyes by use of the process described in Brit. 298,088 (C. A. 23, 280) and Brit. 221,964 (C. A. 19, 1078), as modified or improved upon by the processes described in Brit. 281,536 (C. A. 22, 5536) or Brit. 300,800 (C. A. 23, 5329), by incorporating in the printing paste, together with glycol or a glycol ether, a heavy metallic reserve salt, printing therewith, dyeing and steaming to develop the ester salt, and finally dyeing with the vat dye in the usual manner. Various details and examples are given. Cf. C. A. 24, 5510

Apparatus for dyeing fabrics in open width. FRANK DAVIS Ger 520,000, Sept. 21, 1927. See Brit. 273,207 (C. A. 22, 1800).

Apparatus for dyeing merchandise in the piece. F RANDALL (LITON), LTD. Fr 695,511, May 8, 1933.

Cloth-dyeing machine. MASCHINENFABRIK BENNINGER A-G Ger 516,667, July 21, 1929

Machine for dyeing hank yarns. MASCHINENFABRIK TILM GERBER SÖHNE and GER. WANSLEBEN Ger 520,008, June 22, 1929

Dye vat. RUDOLF THOM Ger 517,732, Oct. 27, 1927. The vat has a drum with rungs on which articles for dyeing are hung.

Tubular heating apparatus for dyeing vats. WILHELM SCHÖNICKER Ger 520,576, Mar. 11, 1930.

Ceramic plates for lining dyeing and bleaching vats, etc. PAUL T GERBER Ger 520,375, Sept. 27, 1929. Structural features are described.

Charts with relatively movable sheets for indicating the correct compositions of different dye baths. I G FARREND A-G Brit. 339,812, Feb. 2, 1929

Dyeing apparatus with telescoping adjustable dye-liquor circulating pipes. C. E. OLDROYD. Brit. 339,009, March 20, 1930. Various structural details are described.

Treating textiles. CAMILLE DREVILLE. Fr 696,224, May 28, 1930. The luster of cloth or textiles of or contg. artificial filaments is diminished by submitting the cloth, etc., to a wearing action by an abrasive on brushes, etc. An app. is described.

Boiling (bucking) textile materials. G ULLMANN Brit. 339,550, Jan. 26, 1929. In obtaining fibers from vegetable materials such as cotton, flax or jute, by removal of fatty and waxy substances, pectin substances, albuminoids, coloring substances and woody and husk portions, using such solvents as hydrousides of Fe or alk. earth or alkali metals, the use of these reagents together is effected and a product capable of continuous bleaching without pressure is obtained by adding to the boiling liquor such materials as allow the metallic soaps of the fats and waxes present to be produced in a sol. or highly dispersive form in which they can easily be rinsed off. For this purpose, various sulfonated preps. and compds. may be added.

Waterproofing textiles. ALEXANDER NATHANSON. Ger 521,029, Sept. 15, 1929. The materials are partly esterified by treatment with a fused fatty acid chloride or anhydride in the absence of a solvent. The excess of esterifying agent is afterward removed by means of a solvent. Thus, a dyed cotton fabric may be immersed for 5 hrs. at 65° in fused stearic anhydride, and the excess of the latter afterward removed with CS₂. Cf. C. A. 25, 1686

Treating textile fibers. WILHELM FRANKER. Fr 696,072, May 24, 1930. Natural or artificial silk is wound on bobbins, etc., in such a way as to form uniform channels throughout so that the treating liquids, gases or vapors flow or are forced under moderate pressure through the bobbins.

Apparatus for the wet treatment of textiles. RENE CLAVEL. Fr. 606,800, May 25, 1930.

Retting fibers. RÖHM & HAAS CO. (SOC. ANON.). Fr. 606,800, May 12, 1930. Vegetable fibers surrounded by ligninous materials are retted in a bath contg. an enzyme, glauher salt and a substance preventing the growth of bacteria such as toluene.

Retting bast fibers such as flax straw. VICKERS, LTD., and O. D. LUCAS. Brit. 339,888, Dec. 23, 1929. The material, after de-seeding, is placed in a tank to which water together with a portion of liquor from a previous first stage of treatment is admitted at a temp. of 25° which is maintained for 12-24 hrs., a portion of the liquor is then run off and may be used for the next batch, and the tank is filled with fresh water and the temp. slowly raised to 37° and this temp. is maintained until retting is completed (after 3-4 days). Various auxiliary details of procedure also are described.

Softening fibers with sulfonic acids. I. G. FARRENTON, A.-G. Brit. 339,888, Aug. 24, 1929. "Cuprammonium silk" or other fibers are softened and rendered flexible by treating with aq. solns. of true sulfonic acids prepd. from carboxylic acids of aliphatic and hydrocarbons contg. more than 6 C atoms, or water-sol. salts of such acids. Use of the mono-Na salt of sulfolalmitic acid or of sulfoheptanoic acid or the salt of tri-ethanolamine with the sulfonic acid or stearic acid is described.

Weighting fibers. OLAV BERG and MAX JEHOFF. Ger. 521,122, Dec. 10, 1925. See Brit. 243,527 (C. A. 22, 504).

Calendering and glazing fabrics. ARNOLD ROSSHAHN. U. S. 1,797,838, March 24. The fabric is dried and electrostatically charged and subjected to heat and pressure (in a roller app. which is described) while in the dry electrostatically charged condition.

Removing stains from fabrics. SOC. ANON. DES PROCEDES R. ARDREBERT. Brit. 340,077, Oct. 23, 1929. Stains are removed from crepe-de-chine or other fabric by the action of the vapor of a volatile solvent in the presence of an adsorbent such as fuller's earth and bleaching charcoal sep'd. from the fabric by a thin cloth or leather. The solvent, such as a small proportion of benzene, may be mixed with the adsorbent, and the treatment may be effected by weighting down the materials between a marble slab and a covering felt.

Fabrics containing natural silk. BRITISH CELANESSE, LTD. Brit. 339,884, March 5, 1929. In the manuf. of fabrics contg. natural silk, especially in combination with yarns of cellulose acetate or other artificial silk, the silk in the state of yarn or fiber is degummed and then sized before weaving, knitting or other textile operations, and after these treatments the size is removed. Dyes are used which can be easily removed with hot water, and various details and examples of materials used are given.

Apparatus for finishing and polishing cloth. WILHELM B. FROMANIER. U. S. 1,797,564, March 24. Structural features.

Centrifuge for the wet treatment and drying of yarns on beams. CARL BAYER. Ger. 519,043, Mar. 12, 1927, and 519,045, Mar. 22, 1928; addn. to 519,045.

Centrifuge for the wet treatment and drying of yarns on beams and bobbins. CARL BAYER. Ger. 519,044, April 18, 1929.

Bleaching animal and vegetable fibers and materials. ÖSTERREICHISCHES CHEMISCHES WERKE G. M. & H. Ger. 516,531, July 6, 1927. Feathers, hair, fur, bristles, etc., are bleached by the action of gaseous H_2O_2 . The materials are kept at about 50°. In the case of furs, etc., the skin is given a protective coating of fat, etc., before bleaching the fur, etc. Cf. C. A. 24, 3672.

Treating wool-washing waters. CARL VON OVERSTRAETEN. Ger. 520,170, Nov. 11, 1927. The waters are treated with a reagent, e. g., CaCl_2 or an acid, capable of pptg. all the suspended or dissolved substances derived from the wool. The ppt. is dried, e. g., by centrifuging, and treated with Na_2CO_3 to convert the Ca soaps or free fatty acids into Na soaps. The product is dissd. with soap soln., and the emulsion so obtained is resolved by centrifuging into sludge, soap soln. and wool fat.

Oiling wool. GASTON L. DUCLOUX. Fr. 604,801, Sept. 18, 1929. Wool is oiled by an emulsion in oil of an infusion of substances such as algae, bakers or gums with the addn. of substances preventing mold formation.

Apparatus for drying wool in air currents. P. MILARD. Brit. 339,907, April 29, 1929. Structural features.

Proofing wool, fur, feathers, hair, etc., against moths, etc. I. G. FARRENTON, A.-G. (Max Weiler and Hermann Stötter, inventors). Ger. 520,184, Apr. 10, 1929. The materials are treated with hydroxybiphenyls or their substitution products contg. halogen atoms or alkyl or alkoxy groups, but not contg. sulfo, carboxy, or nitrogenous groups. Examples are given.

Wool fat. EUGÈNE MERIENS *Ger* 520,008, Oct. 20, 1926. The scum obtained in known manner from wool washing waters is heated under 2-3 atm. pressure without addition of reagents, whereby it is resolved into wool fat and an aq. soln. of soap, etc. An example is given. *Cl. C. A. 24, 4641*

Artificial silk. ACETA G. M. B. H. *Ir* 695,737, May 10, 1930. App. for wetting or oiling threads is described.

Artificial silk cloth. SOC. POUR LA FABRICATION DE LA SOIE "RHODIASETA." *Fr* 696,565, June 10, 1930. Crêpe effects are obtained on cloth composed of or containing strongly twisted cellulose acetate, without sizing the threads before twisting, by treating the cloth in a bath having a very marked swelling action on cellulose acetate, and the peptization action of which is small or nil. 1 sample is given of the use of baths containing TiOH and water, MeOH and water, EtOH and C_2H_5 , and tOH and trichloroethylene.

Desulfurizing and bleaching artificial silk. WILHELM FRANCKE *Fr* 695,845, May 19, 1930. Use is made of baths containing active O , e. g., baths containing H_2O_2 . A desulfurization and preliminary bleaching may take place in a bath already used and a further bleaching in a fresh bath.

Artificial threads, filaments, ribbons, etc. HENRY DRAFFUS *Fr* 695,844, May 19, 1930. The properties of products of cellulose esters or ethers which have been coagulated with aq. agents are improved by the action of heat. The heating may take place in the presence of solvents, swelling or plastifying agents.

Ribbons, bands, films, etc. CAMELLE DRAFFUS *Fr* 695,371, May 9, 1930. A liquid having a limited solvent or softening action on cellulose derivate is applied to a series of filaments of cellulose acetate or other thermoplastic derivate of cellulose so that the filaments are rendered adhesive, and heating, under pressure if necessary, so that they unite to give a product of flattened cross section. The solvent or softening agent may be diacetin, triacetin, dibutyl tartrate or tritolyl phosphate.

Treating regenerated cellulose or vegetable fiber yarns. CHEM. FAB. VORM SANDOS *Ger* 521,121, Oct. 20, 1929. The dyeing properties of yarns of cotton or other vegetable fibers or regenerated cellulose are modified by treating the yarns, while wound on bobbins or the like, in turn with an alkali soln. and a soln. of an esterifying agent, the treating liquids being forced or drawn through the yarns via the hollow cores of the bobbins, etc. Suitable esterifying agents are PbSO_4Cl and its homologs. *Cl. C. A. 24, 5168*

Fireproofing porous materials such as cotton fabrics. HARRY HOPKINSON (to Brunswick Co. of Mass.) *U. S.* 1,797,565, March 24. The material is treated with a soln. of a Pb salt such as Pb acetate and then treated with a soln. of a haloid such as AlCl_3 .

26—PAINTS, VARNISHES AND RESINS

A. H. SABIN

The paint trade—new types of finishes. H. HOULSTON MORGAN, *The Times Trade and Eng. Suppl.* 28, No. 664, 10-11(1931).—Cellulose ester lacquers and synthetic resin varnishes are discussed briefly.

E. M. SYMMES

Covering power of paint. H. LABATE *Peintures, Pigments, Varnis* 7, 1184-92 (1930).

B. HAMILTON

Some considerations on black asphalt paints. MANUEL GONZÁLEZ DE LA VEGA *Rev. Quim.* 6, No. 8, 25-7(1931).—The author shows that paints made with natural asphalts, such as gilsonite, are superior to those having asphalts obtained from petroleum distillate residues. Petroleum asphalt paints under the action of air, sun and rain are turned into a fine dust. This effect is especially noted when the proportion of linseed oil is too great or when Pb or Mn resins are used as driers. A marked difference in weathering properties and fusion points was noted. A natural asphalt resists exposure 3 times as long as a petroleum asphalt. The fusion points were 300° and less than 50° , resp. In a further study of the fusion points, the presence of volatile vehicles altered the results. Exposure for 24 hrs. at room temp. eliminated any volatile matter. Linseed oil lowers the fusion point but after 24 hrs. is converted into knoxyn, which raises the fusion point. The presence of resins and knoleates was disregarded with reference to their effect on the fusion point. The author developed the following method for detg. the nature of asphalts: One-half of the surface of 2 similar glass pieces is painted with a uniform coating. The pieces are dried at room temp. for 24 hrs. A mark is made on the glass close to the painted area and the unknown samples are placed in a

controlled heating device. The heating progresses slowly and the temp. at which the paint runs down and reaches the mark on the glass is found to be significant in each case. The temp. for petroleum asphalt paints was found to be less than 85° , and for natural asphalt paints 135° . The former always run at less than 100° , except in cases of emulsions with asphalts in different phys. conditions.

Titanox-C in paints. M. DOMER. *Paint, Oil & Chem. Rev.* 91, No. 15, 12-4 (1931).—Some of the advantages of Titanox-C are its low sp. gr., its ease of wetting and its non reactivity with acid vehicles. Because of its low sp. gr. it is the slowest settling white pigment available. Other factors involved in the settling properties of paints are (1) non volatile content of the vehicle, (2) character of pigment, (3) mixing and grinding of the paint, (4) use of H_2O , (5) character of the vehicle, (6) use of metallic soaps and methods of incorporating them. With respect to (6) the use of as much as 12% Pb in the oil portion of vehicles is advantageous. To avoid excessive darkening, the Pb is best incorporated on down beats and in the form of basic carbonate.

The viscosity and brushability of paints. HANS WOLFF. *Farben-Ztg.* 36, 1088-9, 1131-2 (1931).—Data are produced to show that the turbidimeter is as suitable for detg. the consistency of paint as the Couette viscometer (cf. C. A. 24, 2313). The plastic nature of paints is brought out more clearly if the data obtained with the latter be plotted as ω against ω instead of simply ϕ vs ω . Curves of viscous materials are parallel to the axis. With either app., the data do not reveal differences in the brushing properties of paints.

Hydrogen-ion concentration and the color of lead chromate pigments. R. C. FRIST, E. PRAGOFF, JR., AND E. L. LITTLEHORN. *Ind. Eng. Chem., Anal. Ed.* 3, 174-6 (1931).—Lead chromates were pptd. from solns. of $Pb(OAc)_2$ and $K_2Cr_2O_7$. The acidity of the solns. was regulated by the addn. of $HIOAc$, HCl , HNO_3 , and H_2SO_4 . H^+ ion concn. was detd. by use of the quinhydrone electrode. Each acid caused the color to change from yellow at pH 3.5 to orange at pH 9, at which there was a max. of brilliance and darkness of shade. For pH values above 9 the colors were off shades. The color changed from yellow at low temps. to a max. at 80° , above which the color again became lighter. Variations in time of pptn. and in concn. of metal ions had no appreciable effect on the color.

Uses of solvents—important applications. THOMAS H. DURRANS. *The Times Trade and Eng. Suppl.* 28, No. 664, 11 (1931).—Nitrocellulose lacquer solvents are discussed briefly.

Glyptals, albertols and coumarone—their value to the varnish manufacturer. E. E. WALKER. *The Times Trade and Eng. Suppl.* 28, No. 664, 7 (1931).—A review. Vinyl resins are also included.

Formula for black, acid-proof stain for wood. F. L. KISKADDEN. *Chemist-Analyst* 20, No. 3, 17 (1931).—Apply 2 coats of hot aq. soln. contg. 4% copperas, 4% blue vitrol and 8% $KMnO_4$. Rub off the excess of the second coat and apply 2 coats of aq. 12% aniline and 18% concd. HCl . When dry apply a coat of linseed oil.

Gum ester. A. NAUROY. *Peintures, Pigments, Vernis* 7, 1214-8, 1262-6 (1930).—A review.

The older compounds—permanently fusible gums and resins. H. COURTNEY BRYSON. *The Times Trade and Eng. Suppl.* 28, No. 664, 17 (1931).—A review of the manuf. of molded plastics.

Phenolic resins—the raw materials. G. T. MORGAN. *The Times Trade and Eng. Suppl.* 28, No. 664, 8 (1931).—A review, with plant illustrations.

Synthetic resins. O. KUULA AND O. ROUTALA. *Acta Chem. Fennica* 4, 1-16 (1931).—A review on the prepn., properties, reactions and uses of synthetic resins.

The constitution of the synthetic resins. GEORG WALTER. *Z. angew. Chem.* 44, 136-7 (1931).—The ability of thiourea to form complex compds. with metal salts, e. g., $CuCl_2$, was utilized in prepg. thiourea-formaldehyde condensation products contg. Cu and Cl . The colloidal solns. of these products gave on electrolyte flocculation, e. g., with KCl , ppts. which, unlike the gelatinous ppts. obtained from urea formaldehyde condensation products (I), may readily be dried to const. wt. The dry ppts. present, because of their content of Cu and Cl , a suitable material for an analysis. Comparative cryoscopic investigations of solns. of methylolthiourea (II) and solns. of the complex compds. prepd. from II and Cu_2Cl_2 , together with the results obtained from the analysis of the flocculation ppts., gave data from which, certain plausible assumptions being considered, numerical expressions were obtained as to the structure and size of the mol. present in the colloidal soln. of I. W. has previously shown (C. A. 21, 3626) that clear

condensation products may be prep'd from the methylol compds. of urea also in the absence of water. It was now det'd that by condensing dimethylolurea in the presence of the generally known condensation agents 6% CH_3O and 22% H_2O were sep'd. These values vary somewhat with the different condensing agents used. In the absence of condensing agents a sepn of 12.5% CH_3O and 15.1% H_2O took place and methylol methylnurea was formed, in agreement with the calcd values. The remaining resin was investigated in each case. The values obtained show that the urea residues must be connected through CH_2 groups (C. A. 23, 2425). W. discusses the structure of the condensation products, the av. no. of links in the chains, etc. Cf. C. A. 22, 4538, 23, 1136, 1123-25, 1398. D. THURSEN

Testing and evaluating stearin pitch (SIEBERT, BLENNEMANN) 27. Furfural and its application in the plastics industry (GENIN) 15. Condensation products containing halogen [starting materials for manufacture of artificial resins] (Fr. pat. 693,602) 18. Aqueous dispersion of TiO_2 (U. S. pat. 1,797,760) 18. Synthetic rubber [products for manufacture of lacquers] (Brit. pat. 340,008) 30. Distillation products of polymeric hydrocarbons (Fr. pat. C96,812) 10.

Paint, Varnish, Lacquer, Enamel and Colour Manufacture. (*New Journal*.) Published monthly by Leonard Hill, Ltd., at 231-2 Strand, London W. C. 2. Vol. 1, No. 1 appeared April, 1931. Price 10s., post free.

HABERT, HANS. Handbuch über die Herstellung und Verwendung der Druckfarben. 2nd ed., revised and enlarged. Berlin: Otto Elsner Verlagsgesellschaft m. b. H. 211 pp.

Handbuch der Lack- und Farne-Industrie. 4th ed., revised. Begun by Franz Seeligmann and Emil Ziehl. Issued by E. Ziehl and Hans Wolff. Edited by HANS WOLFF, FR. ZIMMER and B. SCHEFFLE. Berlin: Union Deutsche Verlagsgesellschaft Zweigniederlassung. 992 pp. M. 45.

MITCHELL, C. AINSWORTH, and HERFORTH, T. C. Inks. Their Composition and Manufacture. 3rd ed., revised. New York: Industrial Book Co., Inc., 336 pp. \$5.

MATTHISON, A. L. Tin Plate Decoration and the Lacquering of Food Containers. Birmingham, Eng.: A. L. Matthison. 156 pp. 12s. 6d.

NOUVEL, OTTO. Die Industrie der Phenol-Aldehyd-Masse. Halle: W. Knapp. 128 pp. M. 9.60, bound, M. 11.

STARKE, ARTHUR. Rechtshandbuch des Lack- und Farnefaches. Unter Berücksichtigung des Chemikalien- und Drogeufaches. Berlin: Union Deutsche Verlagsgesellschaft Zweigniederlassung. M. 9.50.

STERN, ERNST. Farbenbindemittel Farbkörper und Anstrichstoffe mit besonderer Berücksichtigung ihrer kolloidchemischen Grundlagen. Dresden: T. Steinkopff. 95 pp. M. 6.

Taschenbuch für die Farben- und Lackindustrie. 7th ed. Edited by H. WOLFF, W. SCHLICK and W. WAGNER. Stuttgart: Wissenschaftliche Verlagsges. 416 pp. M. 12.50.

WILSON, S.-P. Vernis, émaux, apprêts et mastics de nitrocellulose. Paris: Dunod. 228 pp. F. 50, bound, F. 51.

Paint. JOSEF BLUMENFELD. Ger. 516,748, April 22, 1926. See Brit. 256,302 (C. A. 21, 2902).

Paints. I. G. FARBEWERK A.-G. Fr. 695,523, May 13, 1930. White pigments having good covering power are made by adding to Zn lys a fixed sol. base such as hydrides of alkali or alk. earth metals in excess so that the liquid above the ppt. formed has a pH value of 12 or more.

Testing paint films. KALL FORSCHUNGS-ANSTALT G. m. b. H. (O. F. Kasehitz, Erich Ritter and Paul Nettmann, inventors). Ger. 519,486, Feb. 2, 1930. The method described by Ritter (C. A. 23, 3357) is modified by coating the gelatin layer with a thin layer of metal, concrete or other material to which the paint to be tested is intended to be applied in use. This may be effected by spraying molten metal onto the gelatin layer, or by dusting graphite on to the layer and then depositing metal electrolytically. The paint is then put on the metal, etc., coating and the test proceeded with as before.

Painting with successive coats. E. FRENKEL (trading as the firm of H. Frenkel). Brit. 339,784, Dec. 10, 1928. In a process involving application of successive coats before the underlying material is dry, the binding medium of the coating compns. comprises oil treated with air or gases contg. O_2 so that they are viscous, and a finishing

coat such as a nitrocellulose varnish also may be applied before the preceding coat is dry. Use of various other addnl ingredients also is mentioned.

Driers for use in paints, etc. I G FARBEIND A-G Brit. 339,865, Feb. 22, 1929. Solns of heavy metal or alk. earth naphthenates which are sol in drying oils and org. solvents, for use as driers, are made by dissolving the naphthenates in the fatty acids of drying or semi drying oils such as linseed, wood, soy bean or poppy oils. Examples are given of driers contg Co and Mn.

Driers for paints, etc. I G FARBEIND A-G Brit. 339,922, June 14, 1929. The process described in Brit. 339,865 (preceding abstr.) of prep solns of driers by dissolving heavy metal or alk. earth naphthenates in the fatty acids of drying or semi-drying oils is modified by partially replacing the fatty acids by drying or semi drying oils, e.g., a soln of Pb and Mn naphthenates in various proportions of linoleic acid and linseed oil may be used.

Mixing and grinding machine for paints, enamels, inks, etc. LIONEL G HOLMES Ger 520,495, April 6, 1930.

Pigments. COMPAGNIE LORRAINE DE CHARBONS POUR L'ÉLECTRICITÉ Fr 696,242, Sept 7, 1929. Flocculent pigments of very fine grain are deflocculated by treatment with a dispersing agent composed of an amphoteric colloid in aq soln possessing a suitable elec sign. In the case of pigments composed of lamp black, C black, etc., a protein previously brought to a suitable pH value above 7 is used, the grains of pigment being totally or partially absorbed by the micellas of the protein used. The apparent d of the treated pigment is increased by a final trituration made in the dry state.

Pigments. SOC. ANON. POUR L'IND. CHIM. A BÂLE Swiss 143,293 to 143,295, Aug 21, 1923. Addns to 141,886 (C. A. 25, 2310). Colored pigments which with stand weather, light and water are formed by treating suitable dyes with metals. In 143,293, the azo dye obtained by treating the saponid dye from the diazotized *p*-toluenesulfonic acid ester of 1,8-aminonaphthol-3,6-disulfonic acid and cresidine ($C_{11}H_7OCH_3 \cdot NH_2 = 1,43$) is treated with $COCl_2$ and MgO . In 143,294, the azo dye from the tertiary condensation product (from 1 mol cyanuric chloride, 1 mol of 1,8-aminonaphthol-3,6-disulfonic acid, 1 mol of *p*-aminobenzenesulphonic acid) from 1 mol of $PbNH_2$ and 1 mol of the azo dye from diazotized 1,8-aminonaphthol-3,6-disulfonic acid and cresidine, is treated with pptd chalk or similar Ca compds. In 143,295, the same dye is treated with MgO .

White lead. COMPAGNIA METALURGICA DE MAZARRON (Georg Sitz, inventor). Ger 521,383, Aug. 13, 1924. An aq sludge of $PbCl_2$ or suitable metallurgical products contg $PbCl_2$ is gradually treated, while stirring, with Na_2CO_3 soln, which may contain $NaOH$, the addn. of the Na_2CO_3 soln. being regulated so that the mixt. does not show an alk. reaction until all the Pb has been converted into basic carbonate. Instead of Na_2CO_3 soln, $NaOH$ soln. may be used while passing CO_2 into the mixt. An example is given.

Red lead. H. ERZINGER and CHEMISCHE FABRIK SCHÖNENWERD H. ERZINGER A-G Brit. 340,082, Oct 30, 1929. Red lead of low d is produced by adding to PbO , in an excess of water, small quantities of an org. acid (such as acetic, formic, or amino acids obtained by heating protein refuse) capable of forming basic Pb salts sol in water, treating with CO_2 until a strongly basic Pb carbonate of low sp. gr. and an approx. compn. $4PbCO_3 \cdot Pb(OH)_2$ is formed, sepg this product and oxidizing it to red lead by burning at the lowest possible temp (preferably 100° lower than that usually employed in red lead production, care being taken to prevent pulverization).

Naphthazarine derivatives. I. G. FARBEIND A-G (Georg Kränzlein and Robert Welde, inventors) Ger 520,089, Sept 5, 1929. Addn to 507,347 (C. A. 25, 609). The fastness to light of the products described in Ger 507,347 is improved by treating them with oxidizing agents, e.g., $NaNO_2$, $NaOCl$ or MnO_2 and H_2SO_4 . Examples are given.

Printing inks. HANS ZIMMER. Fr. 696,723, June 6, 1930. The inks are improved by incorporating colors sol in water, particularly tar colors, in the oil base. Colors sol in the oil are also added.

Intaglio printing ink. I G FARBEIND A-G Brit. 339,733, Oct. 23, 1929. An ink suitable for printing on metal foils, celluloid, acetyl cellulose products, cellulose foils, paper, etc., comprises cellulose esters sol in $EtOH$ such as acetylcellulose or nitrocellulose contg. 10.5–11.5% N, coloring matter and a solvent consisting of at least 70% $EtOH$ but free from aliphatic homologs of 1,4-dioxane. Various resins, pigments, oils, fillers, plasticizers, etc., may be added.

Ornamenting surfaces. HELENA S SADTLER. U. S. 1,797,998, March 24. A film such as cellulosic material is secured to a base material and is covered with a layer of

water sol. adhesive such as gum arabic and a protective layer such as varnish or lacquer overlying the base and film bearing a design imprinted on it in coloring matter which would be injuriously affected by the solvent of the protective film.

Thick oils. JOSEF SOMMER. Ger. 517,506, Mar. 27, 1929. Thick oils, such as lithographic or printing varnish are prepd. by heating linseed, etc., oils in a vessel in an atm. of inert gas and subjecting them during heating to tangential currents of superheated steam. App. is described.

Varnish. I. G. FARREND A.-G. (Karl Dämmer, inventor). Ger. 517,902, May 2, 1925. A varnish comprises an aq. emulsion of resin in NH_4 salts of fatty acids, with or without addn. of aq. sol. org. solvents. Thus, ricinoleic acid and colophony are mixed with aq. NH_3 and water. Spirit, methylcyclohexanol, PrOH , etc., and coloring matter may be added. Further examples are given.

Varnishes, etc. BAKELITE CORP. Fr. 695,588, May 14, 1930. Solns. of resins having a basis of thiourea and CH_2O are prepd. by the reaction between thiourea and CH_2O or para-formaldehyde free from water, in the presence of a solvent such as Et lactate. The amt. of hardening agent used is such that a resin of the hardening type is obtained. The reaction may be carried out in the presence of the mono-Et ether of ethyleneglycol.

Lacquers. I. G. FARREND A.-G. Brit. 339,936, Aug. 8, 1929. Lacquers with a base of cellulose deriv. such as nitrocellulose are colored with alk. earth metal salts of azo dyes contg. sulfonate groups obtained by coupling diazotized nitroarylamines with acetoacetic acid arylides. Several examples are given.

Esters of polysaccharide ethers. I. G. FARREND A.-G. Fr. 695,323, May 8, 1930. Cyclic carboxylic acids having more than 6 C atoms are caused to react on polysaccharide ethers with or without catalysts. Thus, benzylcellulose is added to linoleic acid heated to 100° and contg. a little H_3BO_3 . Prolonged heating at 120 – 150° gives the linoleate of benzylcellulose. The products are used in the manuf. of lacquers and plastic masses. Several examples are given.

Coating compositions. Wm. E. SMITH. Fr. 695,922, May 21, 1930. A coating compn. which is resistant to acids and chem. agents consists of approx. pigment 45 and a carrier 55%. The pigment consists of asbestos fiber 28, Al silicate 44 and BaSO_4 28 parts. The carrier may consist of a mixt. of mineral asphalt, steam pitch and petroleum asphalt, melted and reduced with naphtha to produce a quickly drying mixt. resistant to acids.

Oily resinous products suitable for softening celluloid, casein, synthetic resins, etc. I. G. FARREND A.-G. Brit. 339,958, Aug. 12, 1929. Esterification of the hydroxyl groups of aliphatic long-chain carboxylic acids contg. OH groups in the chain, or esters of such acids, with alcs. such as castor oil, ricinoleic acid or dihydroxystearic ethyl ester, is effected by heating with colophony or other natural resinic acid materials, either *in vacuo* or in a current of inert gas, with use of either acid or alk. condensing agents. Several examples are given.

Resin acid derivatives. CHEMISCHE FABRIKEN KURT ALBERT G. M. B. H. Fr. 696,584, June 3, 1930. The acid is transformed into its anhydride and the crude anhydride treated with alcs., phenols, amines, amides, esters, etc. Several examples are given. Cl. C. A. 25, 710.

Resin layers. BAKELITE CORP. Fr. 696,754, June 6, 1930. Layers contg. one or more resinoids, are fixed to a metal or wood support by means of a cement obtained by treating rubber with H_2SO_4 .

Paracoumarone resins. EDWARD H. ELLMS (to Barrett Co.). U. S. 1,797,260, March 24. A paracoumarone resin is made contg. Na sulfonates and Na_2SO_4 , the sulfonate ash being greater than the Na_2SO_4 present as such and the total ash within the range of 0.3 to 1.9%. This product forms stable emulsions in water and oils such as "spindle oil."

Artificial resin coating. COMPAGNIE FRANÇAISE POUR L'EXPLOITATION DES PROCEDES THOMSON HOTSTOV. Fr. 695,951, May 22, 1930. An adhering surface coating forming a thin film is obtained with an alkyl resin reduced to a finely divided state and put in suspension in an inert volatile liquid and appropriately heated after application. Suitable liquids are CCl_4 , kerosene and paraffin hydrocarbons or oil, of the terpene series.

Synthetic resins. AMERICAN CYANAMID CO. Fr. 695,424, May 10, 1930. Synthetic resins are made by heating a mixt. contg. esters of a polyhydric alc. and a polybasic acid, stearic acid or palmitic acid or both and one or more unsatd. fatty acids. An example contains glycerol 100, phthalic anhydride 160, stearic acid 45 and distd. fatty acids from coconut oil 45 parts.

Synthetic resins. **BAKELITE CORP.** Fr 696,390, June 2, 1930 Phenol resins or compns having a basis of phenol resins are made by first making a resin or compn. contg a free phenol, treating the resin with Ca(OH)_2 or an equiv base to transform the phenol into the corresponding phenolate, then mixing with the resinous mass obtained an amt. of $(\text{CH}_3)_3\text{N}_4$ or an aldehyde with equiv hardening action, with or without the addn of appropriate fillers Cf C A 24, 4945

Synthetic resins. **BAKELITE, LTD.** Fr 697,267, June 12, 1930 In the manuf. of resins having a basis of urea and PhOH or their derivs or homologs, and an aldehyde, a solid polymer of the aldehyde such as paraformaldehyde or trioxymethylene is used. The 3 constituents are preferably solid and heated to complete condensation. The urea is in proportion considerably greater than the phenol and the reaction mixt. is maintained neutral or alk. Cf C A 25, 1400

Synthetic resins. **I. G. FARBEN-UND A.-G.** Fr 695,476, May 12, 1930 Neutral resins sol. in oil are made by the reaction of mono- or poly nuclear hydrocarbon derivs, which contain one or more ClCH_2 groups, with aromatic hydrocarbons or their derivs. contg at least one aliphatic or hydroaromatic side chain, or at least one phenolic OH group which may be alkoxylated. Several examples are given.

Artificial resins. **THOMAS & HOCHWALT LABORATORIES, INC.** Brit 340,001, July 10, 1929 See Fr 679,402 (C A 24, 3915)

Synthetic resins. **BRITISH CELANESE, LTD.** Brit 340,101, Nov 17, 1928 BzH is condensed with an aromatic sulfonamide such as benzenesulfonamide or a toluene or xylene sulfonamide, and mixed products may be obtained by including in the reaction mixt. other synthetic resin components such as formaldehyde, furfural, acetone, urea, phenol or aniline. The condensation may be effected without a catalyst or with use of an acid, neutral or alk. catalyst, and the products obtained may be used in cellulose ester or ether compns. for forming lacquers, varnishes, plastic masses, etc., for making films, filaments or the like or as adhesives for manuf. of laminated glass. Pigments, natural resins, etc., may be added. Various details and examples are given.

Synthetic resins. **BRITISH CELANESE, LTD.** Brit. 340,102, Nov 17, 1928 Products are obtained of generally similar character to those described in Brit 340,101 (preceding abstr.) and by a similar process except that furfural is used as an initial material instead of BzH .

27—FATS, FATTY OILS, WAXES AND SOAPS

E. SCHERUBEL

Lecithin retards hydrolysis of fats. **RALF B. TRUSLER.** *Oil and Fat Ind.* 8, 141-3(1931).—In catalytic splitting of fats it was observed that glycerides from various sources behaved differently; animal fats could be split 00% in 10-18 hrs. while crude vegetable oils showed no change for similar treatment. By treating such oils with 5-7% of a 20% soln. of H_2SO_4 or with 1-3% HCl , hydrolysis by means of a catalyst can then be effected. This difference in behavior is explained by the presence of some compd. which inhibits hydrolysis and is removed by the acid treatment. This also explains why a pre-treatment with acid is necessary with the use of the Twitchell reagent. It is thought that the inhibiting substance is lecithin or a closely related substance. By incorporating lecithin with lard and then subjecting the latter to hydrolysis with a catalyst in the presence of oxalic acid for 6 hrs. only 20% of fatty acids was produced. E. S.

A stearin 80 years old. **J. DAVIDSON.** *Seifensieder-Ztg.* 58, 222(1931).—A stearin which had been kept for 80 yrs. in a corked bottle showed the following characteristics: a rancid odor, disappearing when treated with 5% of Frankonite bleaching earth, a pure white color, acid no. 209.5, sapon no. 210.3, I an (Hanus) 7.05, unsapon matter 1.07%; it m. 54° . P. ESCHER

The Twitchell reagent for splitting fats. III. Influence of salt additions and of physical conditions upon the activity of the reagents now on the market and upon the color of the resulting fatty acids. **K. NISHIZAWA AND Y. MATSUKI.** *Chem. Umschau Fette, Oele, Wachse Harze* 38, 73-8(1931).—See C. A. 24, 3915. P. ESCHER

Twitchell reagent. IV. Influence of electrolyte additions on the colloidchemical properties of commercial fat splitters. **K. NISHIZAWA, R. ASADA AND B. SAKUMA.** *Kolloid-Z.* 54, 334-40(1931).—See C. A. 25, 427. V. Various influences on the colloidchemical properties of a new commercial fat splitter, Divulson D. **K. NISHIZAWA AND B. SAKUMA.** *Ibid.* 340-6.—See C. A. 25, 612. ARTHUR FLEISCHER

Testing and evaluating stearin pitch. **GG. SIEBERT AND E. BLENNEMANN.** *Farben-Ztg.* 36, 1090(1931).—The following tests should be applied to stearin pitch: presence of

wool fat or of asphalt, soly, viscosity of solns, drying rate and color, and character of film produced by baking at various temps. G. G. SWARD

Hydrogenation of oils in the presence of a catalyst prepared from nickel borate. L. L. UNBELOHDE AND H. SCHÖNFELD *Allgem. Öl-Fettig.* 27, 425-6 (1930) — A catalyst prepd. by the reduction of Ni borate at 420-470° contained 47.4% B_2O_3 . The free boric acids amounted to about 23%. The reduced Ni borate when cooled in the presence of H_2 was unaffected by atm. O_2 . In the hydrogenation of oils it showed approx. $\frac{1}{2}$ the catalytic activity of a Ni kieselguhr catalyst. On hydrolyzing and washing Ni borate with H_2O it was possible to reduce the B_2O_3 content to 13.5%. This corresponds roughly to the B_2O_3 content of Ni metaborate. Reduction of this hydrolyzed Ni borate at 420-430° resulted in a product having about 3 times the catalytic activity of the originally reduced Ni borate. It was also unaffected by atm. O_2 . It is suggested that it is possible to produce a borate of the same B_2O_3 content by the pptn. of a Ni salt with a mixt. of soda and borax. W. I. BOLLENS

Saponification values of highly colored oils. H. S. JOIS, B. L. MANJUNATH AND S. V. RAO *Half Yearly J. Mysore Univ.* 4, 241-2 (1930) — "The Albert method of detg. the sapon. no. of highly colored oils (cf. H. H. Coburn, *C. A.* 24, 2902) was applied to several oils and found to be very satisfactory." GEORGE CALINGBERT

The determination of inorganic impurities in sulfonated oils. R. HART, *Chem. Umschau Fett, Öle, Wackse Harze* 38, 81-3 (1931) — Referring to Nishizawa-Winokuti's paper (*C. A.* 25, 426) H. discusses the development of his own method now official with the Am. Leather Chemists Assoc. which detg. indirectly the inorg. impurities from the % of ash, the original SO_2 and the alkali combined with soap, and gives formulae of calcn. in the presence and absence of Na_2CO_3 in the ash. When a sulfonated oil is washed with coned. Na_2SO_4 the originally combined H_2SO_4 is more completely neutralized than with NaCl and the oil is then the least soot of all its forms. P. FACHIR

The composition of linseed oils, obtained by different methods and stored under different conditions. K. H. RAUER AND A. FREIBURG *Chem. Umschau Fett, Öle, Wackse Harze* 38, 78-80 (1931) — Three La Plata linseed oils were obtained from one lot of seeds of 1927 by cold pressing, warm pressing and by benzine extrn. Their "consts." did not vary greatly, but in the drying tests the cold pressed oil dried more rapidly and completely than the others, the warm pressed oil being the slowest. The 3 oils were exposed to diffuse daylight and sunlight in well stoppered bottles for 1 yr. and again tested. The drying tests showed the 3 oils in the same order as before, but the total time was very much shortened. The complete analysis by the Eibner-Schmidinger method (*C. A.* 18, 1211) of the fresh oils and the stored oils gave such contradictory results, that no definite conclusions could be drawn, the same is true when they were analyzed by Kaufmann's thiocyanate method and no conclusive relationship could be established. The only definite result is a change in the oils by exposure to light, not measurable by chem. means. P. FACHIR

Thiocyanometric analysis of soy-bean oil. H. P. KAUFMANN *Allgem. Öl-Fettig.* 27, 325-7 (1930) — The fatty acid constituents of oils from various varieties of soy beans were detd. The satd. and unsatd. acids were sep'd. by both the Twitchell and Bertram methods. The individual unsatd. acids were detd. by the thiocyanate method of K. Crude and refined oils from Manchurian soy beans gave, resp.: satd. acids 15.1, 15.7, oleic acid 21.1, 25.7, linoleic acid 54.5, 50.9, linolenic acid 8.1, 7.0, unsaponifiable 1.2, 0.6. Oils from Am. soy beans and so-called Am. "giant beans" gave, resp.: satd. acids 15.0, 14.7, oleic acid 24.55, 24.6, linoleic acid 53.3, 50.2, linolenic acid 6.4, 9.0, unsaponifiable 0.65, 0.57. W. F. BOLLENS

Properties and composition of sandal-seed oil. M. SREENIVASAYA AND N. NARAYANA *Proc. 15th Indian Sci. Congr.* 1928, 160 — Sandal seeds contain 50-55% of a thick viscous oil. On exposure to light, the oil spontaneously thickens to a resinous mass, at 130-40°. S is absorbed and a plastic rubber-like compd. formed. On sapon. with alc. potash a white resinous mass is thrown out which has a very high I no. An ether extd. oil has the following const.: n_D^{20} 1.4790, d_4^{20} 0.9304, I no. 130-140, sapon. no. 185-195, acid no. 20-25, Iiehnert no. 96.97, Reichert-Messel no. 15, Polenske no. 0.2-0.3. The mean mol. wt. of the mixed fatty acids is 288.5. E. J. C.

The hardening of linseed oil at very high hydrogen pressure. H. I. WATERMAN AND J. A. VAN DIJK *Rec. trav. chim.* 50, 279-82 (1931) — In these expts., Ni pptd. on kieselguhr was used as a catalyst (cf. Waterman, *Het. harden van oliën* 5 (1921)), the hardening being carried out in the rocking autoclave described previously (Perquin, *Diss. Delft* 1929) with an initial H₂ pressure of 100-195 atm., and at temps. varying from 35° to 70°. Some of the hardened fats had a pungent odor, strongly resembling that of acrolein but after treatment with water vapor in a vacuum at 150° complete deodorization took

place. For comparison the results of a no. of expts. with the same linseed oil raw material at ordinary pressure in Normann cups at a max. temp. of 180° are also given. A comparison of both sets of figures shows that the hardening at high pressure proceeds in an entirely different manner from that at ordinary pressure, the former especially resulting in the formation of solid fatty acids. The fats hardened to the same I no. contain more solid fatty acids on hardening at high pressure and show a higher m. p., while the SCN I no. is lowered more in the beginning. Consequently at high-pressure hardening a smaller quantity of II is needed to obtain fats of a definite m. p. C. F. VAN DUYN

Recent developments in the fish oil and fish meal industries. H. M. LANGTON. *Food Manuf.* 6, 109-113 (1931). I. II

The chemistry of liquid soaps. H. POMERANZ. *Seifensieder-Ztg.* 58, 230-1 (1931).—The transparency of solid and liquid soaps is due to the soly. of the C_{12} to C_{18} fatty acids (in coconut oil), to hydroxy acid (in castor oil) and to sulfonated acids. P. FISCHER

Washing effect of unfilled soap. C. F. GOMRING. *Dyestuffs* 31, 52-6 (1930).—The theory of detergents is discussed, as well as soap concn., temp., hard water and the germicidal properties of soap. CHAS. E. MULLIN

The influence of lead soaps on the Mackey test. H. WOLF AND R. HEHLINGÖTTER. *Allgem. Öl-Fettztg.* 27, 407-8 (1930).—Oleic acid, free from metallic soaps, after being heated in a Pb container for 20 hrs. at $80-90^{\circ}$ contained 1% of Pb. The same oleic acid, allowed to stand in a Pb container at room temp. for 4 weeks, contained 0.39% Pb. The addn. of 1% of Pb in the form of Pb oleate to an oleic acid not spontaneously combustible in the Mackey app. causes it to become so. The addn. of 1% of Pb to an oleic acid contg. 5% of linseed oil increased its spontaneous combustibility in the Mackey app. W. F. BOLLEYS

Determination of sodium silicate in washing compounds. M. DITTMER. *Seifensieder-Ztg.* 58, 168-9, 185-6 (1931).—D. criticizes the Wüßli standard method which first filters the hot aqueous soln. of the sample, D. decomposes the sample with acid without first filtering, fearing that filtering might remove some SiO_2 which may have become insol. D. further proposes the formula $Na_2Si_4O_9$ for calcn. in place of the present $Na_2Si_2O_5$. P. ESCHER

The Marwa process. O. UHL. *Seifensieder-Ztg.* 58, 167-8 (1931).—The Marwa washing process of E. Walter launders in a washing machine with fatty acids plus an excess of Na_2CO_3 in place of normal soap. U. finds no advantage in this method. P. ESCHER

Evaluation of creolin. P. KARSTEN AND D. VAN OLS. *Pharm. Weekblad* 68, 295-303 (1931).—Methods for the analysis of "creolin," a soap emulsion of crude phenols, are discussed. A. W. DOX

Sulfonated wetting and cleansing agents and related products. CHAS. E. MULLIN AND ROSS M. STRIBLING. *Textile Colorist* 51, 367-71 (1929), *Textile Recorder* 48, No. 571, 65, No. 572, 61, No. 573, 63-5 (1930), cf. *C. A.* 24, 5522.—A review of the patents covering the manuf. and uses of the sulfonated aliphatic oils, including the mineral oils, halogen substituted sulfo acids, etc. CHAS. E. MULLIN

The manufacture and constitution of the new wetting-out and emulsifying agents. CHAS. E. MULLIN AND ROSS M. STRIBLING. *Textile Colorist* 51, 635-6 (1929).—A review of the patents covering the manuf. of wetting, cleansing and emulsifying agents of the aromatic type, such as the compds. of C_{12} to C_{18} mineral oils, ethers, resins, proteins, phenols, etc. CHAS. E. MULLIN

Various applications of the wetting-out agents. CHAS. E. MULLIN AND ROSS M. STRIBLING. *Textile Colorist* 51, 727-30, 807-10 (1929).—A review of the patents covering the use of wetting agents in wetting, scouring, bleaching, mercerizing, dyeing, carbonizing, fulling, spinning, desizing, retting, rayon manuf., polishing and cleaning mixts., dyestuff manuf., inks and paints, absorbents, emulsifiers, etc. CHAS. E. MULLIN

SCHÖNFELD, H.. *Neuere Verfahren zur Raffination von Ölen und Fetten.* Berlin: Allgem. Industrie-Verlag G. m. b. H. 110 pp. M 12

Oils and fats. I. G. FARMENIND A-G. *Fr.* 696,707, June 6, 1930. Oils and fats are freed from mucilaginous substances by heating them with sulfonic acids of org. compds. or their salts. Examples are given of the use of β naphthalenesulfonic acid.

Bleaching fatty acids from marine oils. ARNE GODAL. *Ger.* 520,475, Jan. 27, 1926. The acids are freed from impurities coagulable by H_2SO_4 and are then treated in the dissolved state with up to 5% of concd. H_2SO_4 at $15-20^{\circ}$. Further bleaching with an adsorbent may follow. Cf. *C. A.* 24, 983.

Apparatus for extracting oil from bleaching earths, or for filtering. PHILIPP L. FAUTH. Ger 520,169 Jan 8, 1927. See Brit 283,216 (C. A. 22, 4001)

Apparatus for distilling and recovering solvents used in oil extraction. LOUIS J. SIMON AND SIMON BROTHERS (LACINTRA), LTD. Ger 520,390, June 22, 1927. Brit 278,815 (C. A. 22, 2851)

Solvents and emulsifying agents. MENDEL BURAK. Fr. 697,102, June 5, 1930. Solvents and emulsifying agents for fatty materials and mineral oils are made by adding olein or its equiva to a mixt of soaps or soap equiva such as aromatic or hydroaromatic sulfonic acids or their salts, with ales of high mol wt. Examples are given.

Glycerol. HENKEL & CIE G M B H. Ger 519,470, Oct. 15, 1927. Glycerol recovered from waste soap lye is purified by treatment with aldehydes or ketones in the presence of a small quantity of mineral acid, whereby liquid glycerol aldehyde or ketone condensation products are formed which do not dissolve the saline impurities in the glycerol. After sepp the pptd salts, the condensation products are decompd, e g, by boiling with water. Examples are given. Cf C. A. 24, 5014.

Soap. HENKEL & CIE G M B H. Ger 520,090, Mar 21, 1928. Fatty acid glycerides are saponid by heating them with alkali carbonate soln to about 100° in the presence of a small quantity of a catalytic fat splitter, e g, an alkyl naphthalenesulfonic acid. Emulsifying agents may be included in the mixt. An example is given. Cf C. A. 24, 3670.

Soap. HEINRICH SCIDNIG and ROTHENBURGER SEIFEN- und ÖLAFABRIK SCHMIDT & SCHNEID G M B H. Ger 521,478, Nov 21, 1928. Soap prepd by a cold process is cooled in a machine of the kind used for cooling curd soap. The cooling water is run out of the machine, and the soap soln is then poured in. Cooling water is not supplied to the machine until the evolution of heat due to the reaction has ceased.

Soap. FERRERICO JAACKS-BALLESTER. Swiss 143,405, June 5, 1929. A soap for cleansing and softening the skin even when stained with tar, lubricating oil, etc., consists of water, cacao butter, NaOH, KÖlf and com. C₁₁H₂₃.

Soap. GEORGES M. M. GUYOT. Fr. 696,978, May 17, 1930. A soft soap is made by mixing ordinary black soap with powd. CaCO₃ and adding, with stirring, a satd soln of Na₂CO₃ in repeated amts.

Machine for pressing and cooling soap. SIMON A.-G. Ger. 516,654, April 4, 1928.

Soap-cooling apparatus. KÖNIG FRIEDRICH AUGUST-HÖTTE A.-G. Ger. 517,507, Mar 17, 1927.

28—SUGAR, STARCH AND GUMS

J. K. DALE

Short review of the most important work in the chemistry of carbohydrates for 1929. P. P. SHORUKOIN. Zhurnal Sakharnoi Prom. 4, 413-27(1930)—A review.

Notes on the economic production of sugar. WILLIAM E. CROSS. Rev ind agr. Tucumán 20, 105-14(1929-30)—C. discusses selection of suitable ground, prepn. of the plantation and cane selection, methods of cultivation, fertilizers, watering, harvesting, cane transportation, factory control, grinding, losses in crude juice due to fermentation, defecation, decantation, evapn. crystn., refining, class of sugar most desired, losses, chemicals, fuel economy and by products.

Errors which disturb chemical control of the sugar factory. Wm E. CROSS. Rev ind agr. Tucumán 20, 120-3(1929-30)—C. discusses errors in the detn. of sugar entering, of sugar extd., wt. of sugar cane, water of imbibition, cane fiber, sugar yield, sugar retained, plant efficiency and sugar losses, and he gives an extensive list of the effects of errors on values directly detd. and on calcd. values in chem. control work.

Mechanical refrigeration in the sugar industry. CHARLES DOMBITSKY. Refrigerating Eng. 20, 354-7(1930)—The manus. of sugar from sugar beets is described. Refrigerating processes are of importance in recovering sugar from molasses.

Colloids of a diffusion juice. A. V. DUMANSKII. Zhurnal Sakharnoi Prom. 4, 427-42(1930).

Centrifugal pumps for sugar liquids. M. G. EFIMOV AND I. I. EGOROV. Zhurnal Sakharnoi Prom. 4, 450-8(1930).

The action of sulfides on filter cloth. O. SPENGLER. Deut. Zuckerind. 56, 17-8(1931)—A discussion.

J. P. LEBT

The results of experiments on continuous vacuum pan of the Zuer-Vostokov system. V. I. KOIPAKOV and B. A. BIRDAKOV *Zhurnal Sakharnoi Prom* 4, 458-60 (1930) — The app. for continuous boiling represents a combined vacuum pan and crystallizer. If this app. is compared with vacuum pans there is not much advantage in heat transmission, but there is higher efficiency, decrease in labor and economy of steam.

V. I. BAIKOV

Effect of preliminary liming on the juice work and on the quality of the sugar. I. N. NAUMING *Deut. Zuckerind.* 55, 1353 (1930) — Lab. tests on the Teatini process (C. A. 24, 5529) showed that preliming of diffusion juice with 0.15-0.2% CaO improves the quality of the juice, but that addn. of SO_2 at this stage causes no further improvement. On the basis of this observation, 40 l. of milk of lime was added in the factory in each juice measuring tank, and the prelimed juice passed through the heaters. This gave a juice of yellow color, with a quick settling, dirty green ppt., consisting of large flocs. The alkali was 0.030% CaO. The juice was then further treated as usual, but with only 1.5%, and later 1.25% CaO, instead of the usual 1.75%. With the new method, the color of the juice increased only 1.65° Stammer from the 2nd presses to the 4th effect, while with the old method the increase averaged 5.91°. In factories using only fresh water in the diffusion battery the CaO may be reduced even further to 0.75-1%, instead of 1.25%, with 0.20% for the preliming. The quality of the sugar compared favorably with that obtained by the Teatini process. Other favorable effects of N's process will be reported later.

I. W. ZIRBAY

The effects of sediment during digestion and evaporation of sugar liquors. JOSE HAMON *Lasty Lukroar* 49, 198 (1930). STANEK and PAVLIS (C. A. 23, 4388) showed that if the coned and sand solns. are not filtered before digestion more Ca remained in suspension than if the same liquors were filtered before the digestion. H. carried out his expts. on a com. scale. Unfiltered liquor from the first satn. was passed through a heat interchanger until 96° was attained, pumped into liquors awaiting second satn., satd. during the usual second satn. and filtered. The filtrate was coned at 102° and filtered again. The process involved only 2 filtrations. The above procedure was used throughout the entire 1930 season. Compared to previous seasons, the heat interchanger was free from incrustations (more Ca remained in suspension in unfiltered liquors). One filtration process was saved without interrupting the manuf. or decreasing the quality of the product. Better results were obtained by a single filtration of the coned liquor than from 3 filtrations of liquors before evapn. I. M.

The effect of Atlacide on *Aegimethia indica*. JULIAN A. AGATI and JOHN P. TAN, Bureau of plant industry (Manila, P. I.), *Sugar News* 12, 82-9 (1931) — *Aegimethia indica* on sugar cane. MANUEL L. ROXAS *Ibid.* 89-91 — The application of the weed killer Atlacide in a manner similar to that of fertilizer is effective in destroying *Aegimethia indica*, which is becoming more and more dangerous in the sugar cane regions in the P. I.

V. G. LATA

Filterability of raw cane sugars. I. Effect of various factors prior to pan-boiling. J. C. KRANTZ and H. G. HILL *Ind. Eng. Chem.* 23, 421-4 (1931) — Poor filterability of sirups and sugars is closely assocd. with the quantity of suspended matter. So far no satisfactory method has been found to det. filterability except by actual filtration tests. Pulverized raw sugar shows a much higher filtration rate than the original sugar. The filtration rates of sirup and of the resulting sugar were greatly improved by adding 0.004 g. P_2O_5 to each 100 cc. of raw juice before liming. The total P_2O_5 content of raw juice is no criterion as to whether such juice requires further addn. of P_2O_5 to obtain good clarification, because evidently the P_2O_5 may be only partly in available form. II. Effect of pan boiling operations. *Ibid.* 421-7 — The filterability of sugars is affected not only by that of the sirup from which it is made, but also by pan boiling operations. The principal factors in the latter class are pan circulation, design and manipulation.

F. W. ZERRAN

The removal of some of the soluble aliphatic salts of calcium during saturation. T. NEMES *Z. Zuckerind., Tschoslovak Rep.* 55, 159-62 (1930), *Lasty Lukroar* 49, 315-8 (1931) — Expts. were conducted to det. whether other Ca salts besides Ca formate were adsorbed and removed from soln. by CaCO_3 . A definite quantity of fatty acid (insufficient to exceed the soly. of its Ca salt), together with the corresponding wt. of CaO , was placed in a 200 cc. flask, which was then filled to the mark at room temp. The soln. was heated to 85° and satd. with CO_2 , cooled to room temp. and again made up to the mark. The quantity of H_2O displaced by CaCO_3 was calcd. and this amt. was added. After filtration 50 cc. of the soln. was used for the detn. of the Ca salt as follows: The free lime was titrated with 0.1 N HCl against methyl orange, and the total lime was detd. by pptn. as oxalate and conversion to sulfate. The difference

gave the lime in combination with fatty acid and from this value the adsorbed Ca salt of the fatty acid was calculated. It was found that the adsorption of Ca salts by CaCO_3 was greater with acids of higher mol. wt., e. g., formic, acetic, propionic, isobutyric, caproic and valeric. The results for caproic, which has a higher mol. wt. than valeric (and should, therefore, follow it), cannot be considered accurate because of the impossibility of thorough saturation with CO_2 on account of foaming. J. P. LEBT

Determination of natural alkalinity and calcium salts. I. B. MINTZ AND I. I. SNOIKHET. *Nauk Zapiski Tsukrovo Prom* 10, 131-7(1930) V. E. B.

Optimum alkalinity and calcium salts. S. S. AFANASHEV. *Nauk Zapiski Tsukrovo Prom* 10, 42-50(1930) V. E. B.

Calcium salts in juices and products of beet-sugar fabrication. I. LEVIN. *Nauk Zapiski Tsukrovo Prom* 10, 17-23(1930) V. E. B.

Determination of the sugar losses in the filter-press cakes, following the ordinary and Shapiro's methods. I. I. SNOIKHET. *Nauk Zapiski Tsukrovo Prom* 10, 14-7(1930) V. E. B.

Foam pressure as a measurable quantity. K. SCHMIDT. *Chem.-Ztg* 55, 169-70(1931).—The fact that in the saturation of batches of limed beet sugar syrup with CO_2 , the end point of desired p_{H} is established by observing the height of the foam led to a study of the variations of this height during the process. The total head of liquid and foam at various levels, from the bottom of the vessel to the upper limit of foam, was measured by an apparatus in which the hydrostatic pressure was transmitted through air-vessels to manometers. At the bottom, the head remains constant throughout the process. At the level of the non-foaming liquid, the head rises from zero at the start to a constant value which continues until near the end and then falls to a final value characteristic of the end point in a given vessel. Above this level, positive pressure is observed early in the process (as soon as the foam rises) and follows a curve similar to the above. Thus the process can be controlled by 2 manometers, one at the bottom to indicate the proper amount of charge and one at the quiet level of liquid to indicate the end point pressure. The operation can be made automatic, and the same principle can be applied to a continuous saturation process. J. H. ODELL

Rendering the diffusion juice alkaline when working with dried beet cossettes. O. SPENGLER AND N. LOGINOW. *Z. Ver. deut. Zuckerind* 81, 107-18(1931).—Diffusion experiments were carried out with dried beet chips in a lab. diffusion battery with (1) tap water, (2) water + lime and (3) water + soda. Conclusions: (1a) With tap water the p_{H} falls continuously from the first to the last cell of the battery. (1b) The amount of invert sugar, % sucrose, increases in the course of the diffusion process by about 40%. (2a) On the addition of 0.1% CaO or 0.2% Na_2CO_3 (calculated on fresh beets) equally distributed among the cells of the battery, an almost neutral (p_{H} 6.9) diffusion juice of normal purity is obtained. (2b) The invert sugar, % sucrose, in the juice obtained is practically equal to that in the dried cossettes. (3) The relative viscosity of the juices rendered alk. is higher than for the same juices without alkali addition by some 14% for the 0.1% CaO treatment and by 8% for the 0.2% Na_2CO_3 method. By the addition of alkali to the diffusion battery invert sugar formation is practically completely stopped. The addition of lime to the extent of not over 0.1% on fresh beets is recommended. By the addition of too much alkali the risk is incurred of making the exhausted chips unfit for fodder, e. g., 0.2% CaO gave dirty greenish exhausted chips, while 0.5% Na_2CO_3 gave sticky chips of normal color. F. CAMPS-CAMPS

Calculation of production in raw sugar factories and refineries (beet). J. HAMOUS. *Z. Zuckerind techoslovak Rep* 55, 120(1930).—A discussion with computations. J. F. LEBT

The coagulation of the colloids in the beet juice. A. G. ARKHIPOVICH. *Nauk Zapiski Tsukrovo Prom* 10, 250-6(1930) V. E. B.

The Teatinu process for the clarification of [sugar beet] juices. H. CLAASSEN. *Centr. Zuckerind* 38, 1283(1930).—A discussion. J. P. LEBT

The mathematical expression of the composition of sugar-beet juice and other questions of beet analysis. F. KAVL. *Z. Zuckerind techoslovak Rep* 55, 60-3(1930).—A discussion. J. F. LEBT

Observations on over-saturation and saturation [of beet-sugar juices]. BREYER. *Deut. Zuckerind* 55, 1201-2(1930).—A discussion. J. F. LEBT

The use of sulfur dioxide in the purification of sugar [beet] juices. K. SOLON. *Deut. Zuckerind* 56, 119-20(1931).—A discussion with references. J. F. LEBT

The return of diffusion and press waters to the process in the beet-sugar factory. P. HIRSCHFELDER. *Centr. Zuckerind* 39, 12-3(1931).—A discussion. J. F. LEBT

The yield of dry [sugar beet] pulp and the specific gravity of moist and dry pulp.

F. KRYŽ Z *Zuckerind. czechoslovak. Rep* 55, 128-9(1930). *Listy Cukrovar* 49, 284-5 (1931) J. F. LEFTE

Hydrolysis of pulps. M. S. FILOSOFOV. *Nauk. Zapiski Tsukrovoi Prom* 10, 419-22(1930).—In order to det. the sugar losses in pulps a hydrolytic method was tried. The pulps were treated with concd H_2SO_4 , then dild to 0.7 N and boiled for 2 hrs. From 45.62 to 62.46% (on the wt of dry substance) sugar was obtained by the iodometric method. On fermentation 18.76% of alc. on the wt. of dry substances (minus the alc. from sucrose) was obtained. The nature of the sugar was not investigated. V. F. BAIKOV

Contribution to the knowledge of crystallizer operation. P. HONIG AND W. F. ALEWYN. *Sugar News* 11, 727-34(1930). 12, 22-34, 97-100(1931).—See C. A. 22, 1698. V. G. LAVA

Annual analytical report of the research department for starch manufacture. A. PARLOW. *Z. Spiritusind* 53, 311(1930) S. J.

Practical yield tables for dried potatoes. B. LAMPE. *Z. Spiritusind* 53, 326-8 (1930).—L. has revised the tables of Parlow. *Ibid* 347-8.—The use of the tables is given. S. J.

The constitution of starch. M. SCHÖEN. *Bull. soc. chim. biol* 12, 1033-99(1930), cf. C. A. 24, 4951.—A review with bibliography. C. G. K.

The chemical study of chicory and topinambur in connection with the problem of mulin and crystallized fructose production. G. L. ENHORN, A. B. MILSKII AND E. YA. KALASINTSEV. *Nauk. Zapiski Tsukrovoi Prom* 10, 143-50(1930) V. E. B.

The effect of the soil reaction on beet yields (JORET) 15. Losses of water-soluble H_2PO_4 by the clarification of molasses under acid conditions and heat (HUMMER) 16. Beet infections with *Cercospora beticola* Sacc. (CHYZANOWSKI) 15. Molasses as a fuel (RODZIEWICZ) 21. Vitamins in sugar-cane juice and in some cane juice products (NELSON, JONES) 11E. Tartaric compounds [from vinasses] (Fr. pat. 697,179) 16. Automatically controlling [bleaching and filtering of sugar solutions] (U. S. reissue 18,005) 13. Electrically heated apparatus for concentrating sugar juices (Brit. pat. 339,649) 4. The use of denatured sugar, molasses-fodder and sugared cosettes in agriculture (ŠANDERA) 12.

Methods of Chemical Control for Cane Sugar Factories of the Association of Hawaiian Sugar Technologists. 6th ed., revised. Honolulu: Advertiser Pub. Co. 140 pp.

Apparatus for drying and transporting beet roots. SUGAR BEET & CROP DRIERS. LTD. Fr. 697,283, June 12, 1930.

Crystallization apparatus for sugar, etc., with rotary heating or cooling tubes. THEODORE GRUENWALD. Ger. 521,104, April 20, 1928.

Treating sludge from sugar manufacture, etc. A. BORSIG G. M. & H. Ger. 520,298, Apr. 18, 1928. Sludges such as are obtained in washing beets are led to a rising column of water, the velocity of which is regulated so that clean sand sinks to the bottom of the column.

Zinc calcium formaldehyde sulfoxylate. EDGAR BOULOGNE and SOC. INDUSTRIELLE DES DÉRIVÉS DU SOUFRE. Ger. 521,204, Nov. 27, 1929. An insol. double salt is prepd. by warming a mixt. of Zn formaldehyde bisulfite and neutral Zn formaldehyde sulfoxylate with a sol. or insol. compd. of Ca, e. g., $CaCl_2$, milk of lime or bleaching powder. The product may be used in sugar manuf. and in printing textiles. Examples are given.

Starch. ERNST STERN. Ger. 519,300, Aug. 6, 1924. A product that swells in cold water is prepd. by treating an aq. soln. of alkali starch at atm. temp. with a salt of an alk. earth or heavy metal, e. g., with $BaCl_2$, sepg. the resulting ppt., and mixing the ppt., after drying it if desired, with a water-sol. alkali salt. Examples are given.

29—LEATHER AND GLUE

ALLEN ROGERS

A new and improved method of moisture determination and its application to leather. A. COLIN-RUSS. *J. Intern. Soc. Leather Trades' Chem.* 15, 113-26, 166-82 (1931).—An app. is described wherein leather is heated in a tube surrounded by vapor

of a boiling liquid, and the H_2O evolved from the leather is allowed to react with $CaCl_2$ forming CaH_2 which is collected over Hg and measured. A blank on the $CaCl_2$ is indispensable. For standardizing the procedure, numerous hydrates and liquid H_2O were used, the latter, as well as Glauber's salt, give up their H_2O so rapidly that the reaction with $CaCl_2$ is incomplete, borax and gypsum are not completely decomposed at 100° . $(NH_4)_2C_2O_4 \cdot H_2O$ gave the theoretical yield of H_2O , and the time gas volume curve is that of a monomolecular reaction. With leather of heavy vegetable tannage, it was found that the percentage water obtained by the method is almost independent of the reaction temp. (controlled by suitable choice of surrounding solvent vapor) over the range from 61.5° to 133.5° : 12.6% at 61.5° , 13.1% at 76.5° , 13.8 at 100.5° , 13.05 at 133.5° . Time required to complete the reaction decreased as temp increased. Degree of subdivision of the sample is practically without effect. In a sumac tanned leather, which showed marked contraction at 100° , percentage H_2O obtained by the method decreased from 13.9% at $61-71^\circ$ to 10.3-11.0% at and above 100° , the decrease indicating that the contraction involves fixation of H_2O . Comparative results on 4 leathers by the carbide and by the oven drying methods showed that the former always gives lower results the difference varying from 1.45% for Cr tanned calf to 6.03% for heavily tanned vegetable cowhide. Good agreement was obtained between results yielded by the carbide method and those obtained by very prolonged desiccation over H_2SO_4 , except in the case of the abnormal sumac leather. H. B. MERRILL

Salt stain in the manufacture of leather. P. WHITE AND F. G. CAUGHLEY. *New Zealand J. Sci. Tech.* 12, 108-113 (1930).—Lab. scale tests show that the growth of the halophilic organism responsible for "red heat" of hides is inhibited in media and on skin by adding 0.5-1.0% NaF or $NaSiF_6$ to fresh curing salt. With used curing salt, the fluorides were not effective. Fluorides appear to cause unusual shrinkage of hide during curing, the effect largely disappears after liming. Practical tests are projected. H. B. MERRILL

Deliming and bating. E. STASNY. *Schweiz. Leder- und Ztg.* No. 25 (1929); *J. Am. Leather Chem. Assoc.* 25, 46.—The theory of deliming and bating is discussed. The most important role of trypsin is believed to be that of producing a partial peptization of the collagen fibers. This action is more marked in the presence of chlorides than in the presence of sulfates, for this reason NH_4Cl is preferred to $(NH_4)_2SO_4$. The activity of the bate falls off rather slowly on the alk. side of the optimum pH range (which is from 7.5 to 8.5), being little less at $pH = 10$ than at 8.5, at pH values lower than 7.5 the decrease in activity is more marked. H. B. MERRILL

High tannin yield from South African gum trees. G. SMITH. *Shoe and Leather Reporter* 179, No. 6, 16 (1930); *J. Am. Leather Chem. Assoc.* 25, 427.—The bark of *Eucalyptus asirgens* has been found to contain 33-44% tannin. The rate of growth is rapid enough to justify com. culture. H. B. MERRILL

Theory of chrome tanning. K. H. GUSTAVSON. *J. Tech. Assoc. Fur Ind.* 2, 7-26 (1931).—A review. H. B. MERRILL

Camphor as a preservative for tan liquors. P. D. DALY. *Proc. 15th Indian Sci. Cong.* 1928, 154.—*Aspergillus niger* and certain other fungi constantly found in tan liquors bring about a reduction in the tanning value of the liquors because of the fermentation of the tannic acid induced by the organism. The action of various disinfectants or preservatives in diminishing this loss has been investigated, and of those tried camphor seems the most promising, since it is efficient in small quantities and without harmful effect on the activity of the liquor. E. J. C.

The washing and utilization of chrome leather shavings. E. SAUER AND W. ESCHMANN. *Kolloid Z.* 54, 326-34 (1931).—The raw material showed the following analysis: H_2O 14.8, ash 10.7, Cr_2O_3 5.9 and N 11.8%. Tests on the extn. of Cr_2O_3 were made at 20° and 30° with 0.5-10% H_2SO_4 , treatment with 1% alkali before H_2SO_4 treatment, treatment with lime and subsequently HCl , and treatment with dil. H_2O - $NaOH$ soln. The highest extns. (93%) were with pretreatment with $NaOH$ and subsequent H_2SO_4 extn. and with the H_2O - $NaOH$ solvent. A dark-colored glue may be extd. from the residue. A. F.

Level dyeing in leather. WALTER C. DURFER. *Dyestuffs* 31, 125-7 (1930).

Glues and gelatins—varied applications. J. C. KERNOT. *The Times Trade and Eng. Suppl.* 28, No. 664, 24 (1931).—Examples of uses are given. E. M. S.

Condensation products containing halogen [starting materials for manufacture of tanning materials] (Fr. pat. 695,602) 18.

GRASSER, GEORG *Kurzes Lehrbuch der Chromgerbung. Ein Leitfaden für Praktiker und Theoretiker* Stuttgart F Enke About 210 pp. About, M. 14; linen, M. 16

Leather. ALBERT J DUCAMP and CHARLES F A GUYNOT *Fr* 696,252, Sept 9, 1929 The hardness and resistance to wear of leather is increased by introducing into the leather chem substances having the character of anti-catalysts such as cyanides, particularly $\text{Hg}(\text{CN})_2$, or organometallic compds such as PbI_2 , PbI_2Ph , and Fe carbonyl , or phenols such as resorcinol or their derivs The catalysts are used in soln in appropriate solvents

Stretching and drying skins, hides, etc. THE LEATHER MAKERS PROCESS CO *Ger* 516,568, *Feb* 18, 1930

Stretching and drying hides. MACHINERY DEVELOPMENT CO *Ger* 516,569, *Jan* 8, 1929

Tanning. CHARLES L. MAVER *Fr* 696,254, Sept 10, 1929 Acrolein in the form of vapor or in soln in animal, vegetable or fish oils is used for tanning leather, etc *Fr* 696,255 describes the tanning of leathers by treating them with phenolic compds (phenols cresols, naphthols etc) and aldehydic compds ($\text{C}_6\text{H}_5\text{CHO}$, AcH , acrolein, etc.) in succession The products may be used in soln in animal or vegetable oils or fats

Tanning. OTTO RÖHM *Ger* 520,091 *June* 21, 1927 H_3PO_4 or H_2AsO_4 or salts or esters of these acids are added to mineral tanning baths Examples are given

Tanning agent. OSKAR A MULLER *Ger* 519,267, *Feb* 10, 1929 Lignous cellulose materials are chlorinated while moist and then extd with a water sol org solvent, e g EtOH , in the presence of a strong mineral acid The ext is freed from chlorohydrin by pptg thus with water, and the org solvent is removed by distn The tanning agent can then be extd from the aq residue with AcOH , or pptd therefrom with Ba(OH)_2

Tanning and cleaning agents. RÖHM & HAAS CO *Fr* 696,327, *May* 30, 1930 Compds formed by condensing a phenol with an aldehyde in the presence of a small quantity of acid as catalyst are afterward condensed at high temps with an unsatd fatty acid having 10 or more atoms of C, such as oleic acid, and the resin obtained is sulfonated to obtain a product sol in water The products are useful for tanning, cleaning, emulsifying, wetting, mordanting, as penetrating agents for acid coloring materials and as dispersing agents for dyes

Detanning chrome-tanned leather. HENKEL & KRAUSE, CHEM FAB G M B H (Frich Krause, inventor) *Ger* 521,477, *Nov* 11, 1928 The leather is treated with alk. reacting borates, or with alk reagents in the presence of sol borates The Cr compds are then extd with acid. Examples are given

Dressing for leather. C I DU POY DE NEMOURS & CO *Fr* 696,523, *May* 23, 1930 A dressing for leather contains a cellulose deriv such as the nitrate or acetate, a wax, a softening agent, a solvent and a diluent. An example contains cellulose nitrate of high viscosity 17, dibutylphthalate 08, carnauba wax 17, TiO_2 3.3, AcOEt 15.5, AcOBu 10.3 and alc 66.7% Other examples are given

Preserving bone material. BRITISH GLUES & CHEMICALS, LTD, and R B DREW *Brit* 340,010, *Sept* 7, 1929 Fresh bones are subdivided as by cutting into shavings, followed by drying, and, before drying, may be extd with a solvent such as water contg NaCl to remove blood, fat-splitting enzymes and other undesirable constituents An app and various details of operation are described

Transparent gelatin, etc., layers. GREGGER KLOTZ *Ger* 516,580, *Feb* 6, 1926 Transparent ions of gelatin are toughened by coating with transparent ruhler

Working up animal wastes, etc. GUSTAV HÖNNICKS *Ger* 516,810, *Dec* 28, 1928 The waste is treated with super heated steam to ext sol matter and glue The ext is treated with a pptg agent to ppt the extd matter and the glue, which may be then added to the solid residue remaining after extn

30—RUBBER AND ALLIED SUBSTANCES

C. C. DAVIS

The plasticizing of rubber. F. JACOBS *Rev. gén caoutchouc* 8, No. 69, 17-27 (1931), cf *C. A.* 25, 1705—A review and discussion of the properties of gutta resins, various pine products and fatty acids.

Hard rubber—its main applications. T. R. DAWSON. *The Times Trade and Eng.* *Suppl.* 28, No. 664, 18(1931).

C. C. DAVIS
R. M. SYMONS

Organic rubber colors. F. HARRIS COTTON *India Rubber J.* 81, 413-4 (1931); cf *C. A.* 25, 1117—A general discussion, with 21 references. C. C. DAVIS

Stretch in rubber transmission belting. C. W. STAACHE. *Proc. Am. Soc. Testing Materials* 30, Pt. II, 941-56 (1930)—The expts. were made to det. the proper inelastic stretch to remove from belting contg. a given fabric during the vulcanization process. C. C. DAVIS

A study of the performance characteristics of a 4-inch 4-ply rubber transmission belt. J. E. SKANE. *Proc. Am. Soc. Testing Materials* 30, Pt. II, 928-43 (1930)—The tests were made to det. the performance of the belt under varying loads, speeds and slips. C. C. DAVIS

Substitutes for natural rubber—difficulties of synthetic process. HENRY P. STEVENS. *The Times Trade and Eng. Suppl.* 28, No. 664, 26 (1931).—A review. E. M. SYMMES

Synthetic rubber. ERNST KLEINER. *Kunststoffe* 20, 5-6 (1930)—A review. B. II

Distillation products of polymeric hydrocarbons (Fr. pat. 606,812) 10.

Rubber. ALBERT C. BURRAGE, JR. Fr. 697,090, June 4, 1930. Oxidation or aging of vulcanized rubber is prevented or reduced by substituted nitrosoamines such as diphenylnitrosoamine, phenyltolyl nitrosoamine, di- α -naphthyl nitrosoamine, etc.

Porous rubber. K. D. P., LTD. Ger. 621,307, Mar. 31, 1928. Latex is converted by mech. means into foam, if desired with the aid of a foaming agent, and the foam is vulcanized. Thus, a mixt. of latex, S, ZnO and an accelerator may be agglomerated with (HCOO)₂Ca soln. and the mass beaten to a stable foam after addn. of saponin. The foam is charged into molds and vulcanized.

Rubber latex. I. G. FARBENIND A.-G. Ger. 510,483, Nov. 12, 1926. Salts of true sulfonic acids having good wetting properties are used as preservatives for rubber latex, alone or with other preservatives. The salts may be used also as assistants in the impregnation of fabrics, cork, paper, etc., with latex. Examples are given.

Coagulating latex. METALLGES. A.-G. Fr. 695,786, May 17, 1930. Latex is coagulated by adding substances which do not on addn. have a thickening or coagulating action, but which on a change of phys. conditions such as change of temp. introduce substances which have a coagulating action in an amt. sufficient for the coagulation of the latex. A complex salt which gives a coagulating ion on dissociation by heat, a salt which dissociates on heating to give an acid, or a coagulant protected by a layer of inert adsorption substance which is removed by heat may be used.

Rubber inner tubes from latex. EDWIN B. NEWTON (to American Anode, Inc.). U. S. 1,797,240, March 24. Various details are described of uniting valve pads of masticated rubber to wet coagulated tubes, pressing, drying and vulcanizing.

Electrodeposition of rubber, etc. THE ANODE RUBBER CO. Ger. 520,323, April 30, 1926. See Brit. 253,085 (*C. A.* 21, 2138).

Preserving rubber. MARION C. REED (to B. F. Goodrich Co.). U. S. 1,797,241, March 24. An addn. product of an aromatic nitro compd. such as dinitrochlorobenzene with a secondary aromatic amine such as phenyl β -naphthylamine is added to rubber compns. as a preservative (suitable in the proportion of about 0.5%). Various other examples are given also. *Cf C. A.* 25, 2331.

Rubber preservation against aging. P. C. JONES (to B. F. Goodrich Co.). Brit. 339,834, Jan. 21, 1929. Aging is retarded by adding tetrasubstituted hydrazines (several of which are mentioned as suitable) to latex, to rubber before vulcanization or to the surface of rubber after vulcanization. Some description of the manuf. of these preservative substances also is given. *Cf C. A.* 25, 017.

Rubber dispersion. WALDO L. SEMON and RICHARD A. CRAWFORD (to B. F. Goodrich Co.). U. S. 1,797,243, March 24. A mixt. of rubber 100, benzene 10-100, glue 1 and casein 1 part is masticated while slowly adding an aq. soln. of K oleate until the rubber constitutes the dispersed phase of the batch. The product is suitable for use as a cementing medium. *Cf C. A.* 24, 2918.

Filtering apparatus suitable for filtering rubber dispersions. ANDREW SZBOVARI and CHARLES M. SPENCER (to American Anode Inc.). U. S. 1,797,248, March 24. Structural features of an app. having a filtering material sub. as a textile fabric strip.

Apparatus for making spongy articles from organic dispersions. THE DUNLOP RUBBER CO., LTD., and THE ANODE RUBBER CO., LTD. Fr. 696,176, May 27, 1930.

Coloring rubber, etc. I. G. FARBENIND. A.-G. (Arwin Ranft, inventor). Ger. 506,207, Nov. 14, 1924. Addn. to 462,221. Natural or synthetic rubber, gutta percha

or balata is colored with water insol. vat, sulfur or azo dyes or lakes by incorporating in the rubber, etc., an aq. soln. of the leuco compd. or other sol. modification of the dye, or an aq. soln. of the dye or lake components, and then converting the leuco compd., etc., into the insol. dye or lake in known manner. Examples are given. Cf. C. A. 25, 2021, 2331.

Light rubber board. DUNLOP RUBBER CO., LTD., and E. W. MADGE. Brit. 310,024, Sept. 19, 1929. Light rubber board suitable for use in airplane and speedboat construction comprises one or more porous layers of hard rubber (obtained from aq. dispersions) sepg. 2 or more layers of metal, ebonite, ebonite-coated metal, fabric-reinforced ebonite or ebonite-impregnated fabric or "doped fabric" or plywood. The aq. dispersions used may contain rubber, gutta-percha, balata and similar vegetable resins, natural or artificial, vulcanized or unvulcanized, etc. Various details of manuf. are described.

Forming rubber tubes suitable for use as tire inner tubes. ERNEST HOPKINSON and WILLIS A. GRINNOVS (to Morgan & Wright). U. S. 1,707,580, March 21. An aq. dispersion of rubber is placed in contact with a surface of a tubular form of fibrous material such as woven fabric and fluid from the dispersion is withdrawn through the form to deposit solids from the dispersion in tubular shape on the form, the thickness of the deposit adjacent to its ends is gradually reduced to form skived ends, the deposit is dried and removed, and its skived ends are united. App. is described, as are also various compns. used.

Rubber threads. DUNLOP RUBBER CO., LTD., W. G. GORHAM and E. A. MURPHY. Brit. 339,676, Sept. 13, 1929. Threads are made by cutting an unvulcanized sheet which may be produced directly from latex, and vulcanization is then effected (after skiving, if desired). The sheets may be made, from a dispersion prepd. as described in Brit. 290,313 (C. A. 23, 1012), by the method described in Brit. 302,201 (C. A. 23, 4375).

Stiffening and impregnating fibrous materials such as "shoe socks." DUNLOP RUBBER CO., LTD., D. F. TWEISS and E. A. MURPHY. Brit. 339,974, June 18, 1929. Shaped cellulose or other fibrous products such as shoe socks are treated with the foam obtained by whipping up aq. dispersions such as those of rubber which may be prepd. as described in Brit. 332,525 and Brit. 332,526 (C. A. 25, 437), and the treated material is cured and dried or subjected to a setting treatment as described in Brit. 303,544 (C. A. 23, 4596). The material may be preliminarily stiffened, waterproofed and coated with acid latex.

Rubber impregnation of ropes, cords, etc. D. P. FROST and BRITISH ROPE CO., LTD. Brit. 340,051, Oct. 3, 1929. Some of the yarns to be formed into a strand are treated with an aq. rubber dispersion in such quantity that when the treated yarns, together with untreated yarns, are formed into a strand, the strand will have the desired degree of impregnation. Various details of procedure are described.

Apparatus for manufacture of cords or strings of rubber-impregnated materials, etc. REGINALD TRUESDALE, ROBERT C. SMITH and EDWARD SIMPSON (to Dunlop Rubber Co., Ltd.). U. S. 1,797,249, March 24. Structural features.

Putting rubber tags on laces. A. SCHÖELER. Brit. 340,173, Feb. 27, 1930. A mixt. of coned latex and S is applied at the end of a bollow lace, which is pressed into a needle form in a heated mold (the mixt. being transformed into soft rubber). To accelerate vulcanization, a metal salt or oxide in alk. soln. may be added, with other materials such as tannin, phenol, lampblack, ochre, etc. Various modifications of the procedure also are described.

Pencil eraser. STYLLABEMENTS P. OROONGE ET CIE. Fr. 695,258, May 7, 1930. A rubber suitable for removing pencil marks, etc., contains starchy materials, a suitable compn. being natural rubber 4, starchy substances 10, petrolatum 4, vulcanized waste 2, abrasives 2, faetice 1, lithopone 3, S 0.1 and accelerators 0.05 part. The mixt. may be vulcanized or not.

Polymerizing butadiene hydrocarbons. I. G. FARBENIND. A.-G. Fr. 695,299, May 8, 1930. The polymerization of butadiene hydrocarbons using alkali or alk. earth metals is carried out in the presence of small quantities of unsatd. ethers such as vinyl ethyl, vinyl butyl, allyl ethyl or propargyl ethyl ether. Several examples are given. Cf. C. A. 25, 1412.

Polymerizing diolefins. I. G. FARBENIND. A.-G. Fr. 695,441, May 12, 1930. In the polymerization of diolefins by alkali or alk. earth metals, the reaction is carried out in the presence of org. compds. contg. an atom of C, 2 valencies at least of which are satd. with O. Acetals, especially cyclic acetals and among them those contg. a double bond, are particularly advantageous. Several examples are given.

Diolefin polymerization products. I. G. FARREND A-G (Georg Ebert and Friedrich A. Imes inventors) Ger. 520,104, Jan. 26, 1929. The polymerization of butadiene and other diolefins by means of alkali or alk. earth metals is improved and accelerated by addition of a small quantity of a cyclic ether, e. g., dioxane. The products may be made into filaments or films, which may be hardened by heat or by vulcanization. The vulcanized products resemble soft rubber. Examples are given.

Artificial rubber-like masses. I. G. FARREND A-G. Fr. 695,300, May 8, 1929. Polyvinyl alcohols are condensed with aldehydes in the presence of catalysts such as AlCl_3 , ZnCl_2 , or NaHSO_4 . The products may be vulcanized to hard substances resembling ebonite.

Synthetic rubber. I. G. FARREND A-G. Fr. 695,745, May 16, 1929. Masses resembling mixed rubber are made by taking polymerizates, finished or unfinished, of diolefin hydrocarbons and continuing their polymerization with one or more analogous or homologous diolefin hydrocarbons under the same or modified conditions. Examples are given.

Synthetic rubber. I. G. FARREND A-G. Fr. 696,147, May 27, 1929. Diolefins are polymerized in the presence of alkali or alk. earth metals, their mixts. or alloys in the form of particles of uniform size.

Synthetic rubber. I. G. FARREND A-G. Brit. 347,008, Aug. 19, 1929. Polymerization of diolefins with alkali or alk. earth metals is effected in the presence of an org. compd. contg. a C atom with at least 2 valencies satisfied by O, such as satd and unsatd. acyclic or, preferably, cyclic acetals, or the components of such acetals, e. g., dibutyl acetal, ethylene acetal or other acetals from 1,2- and 1,3-glycols with aldehydes, unsatd. acetals such as those from crotonaldehyde or acrolein and butylene glycol. Various aliphatic, aromatic or hydroaromatic ketones also may be added, as may aldehydes, esters, small quantities of acids such as HOAc or formic acid, etc., and either tough or soft and plastic products can be obtained by varying the proportions of the added substances. The products are suitable for vulcanizing or for the manu. of coatings, films, lacquers, artificial silk, etc. Numerous details and modifications of procedure are described.

Synthetic rubber, etc. I. G. FARREND A-G. Brit. 347,004, Aug. 12, 1929. Unsatd. org. compds. such as acrylic, butadiene, isoprene and other long-chain acids or their derivs. such as their esters and their mixts. (e. g., crude linseed oil), unsatd. aliphatic hydrocarbons with more than one double linkage such as butadiene, isoprene or other diolefins or their mixts., or unsatd. aromatic compds. such as styrene are polymerized in the presence of heavy metal carbonyls such as those of Fe, Ni, Co, Mo, W or Cr. Numerous details and modifications of procedure are described.

Storing unvulcanized rubber sheets. P. BERRY (to Goodyear Tire & Rubber Co.). Brit. 339,720, Feb. 26, 1929. Adhesion of the sheets is prevented by interposing between their surfaces a non-adhesive liner comprising fabric which has been treated with soap soln. (suitably soap prepd. from coconut oil).

Transparent vulcanized rubber. DOUGLAS P. TWISS and EDWARD A. MURPHY (to Dunlop Rubber Co., Ltd.) U. S. 1,776,250, March 24. Deposits or products are formed directly from a mixt. of rubber latex, S and an ultra-accelerator such as Zn piperidine carbothionolate without use of ZnO , and vulcanized.

Rubber vulcanization accelerators. DUNLOP RUBBER CO., LTD., D. P. TWISS and F. A. JONES. Brit. 347,063, Nov. 1, 1929. Accelerators are prepd. by the interaction of halogen derivs. of C_4H_8 or its homologs with dithiocarbamates. An example is given involving the use of piperidine piperidine 1-carbothionolate and 4-chloro-1,3-dinitrobenzene, and other suitable starting materials also include Zn diethyldithiocarbamate, diethylammonium diethylthiocarbamate and picryl chloride. Various details of procedure are described.

Rubber vulcanization. DEUTSCHE HYDRIERWERKE A-G. Brit. 339,826, Jan. 8, 1929. Vulcanization is accelerated by the addn. of esters of cyclic alcoh. such as the homyl ester of adipic acid, cyclohexyl esters of hydrophthalic acid, and cyclohexyl ester of benzoic acid.

Vulcanizing synthetic rubber. I. G. FARREND A-G. Fr. 695,269, May 7, 1929. See Brit. 335,970 (C. A. 25, 2722).

CHEMICAL ABSTRACTS

Vol. 25.

JUNE 20, 1931

No. 12

1—APPARATUS AND PLANT EQUIPMENT

W. L. BADGER

Two laboratory devices. M. TROMER. *Bull. soc. chim.* [4], 49, 185-6(1931).—In a modification of the ordinary coil gas-washing bottle the upward movement of the bubbles is utilized for producing a continuous circulation of the washing medium through the 2 limbs of the app. An open vessel of any size can be used for holding the washing medium. In an app. (of glass) for producing hydraulically a high-pressure air blast the water level, which det. the air pressure, is automatically maintained for any set water pressure by means of a float valve. Sketches are included. W. A. LA LANDE, JR.

Laboratory apparatus for tube distillation. HERGO BURSTIN AND JOSEF WINKLER. *Chem. Ztg.* 55, 212(1931).—This is an illustrated description of a continuous vacuum distn. outfit of cast iron, Cr-Ni steel or other metal which was originally designed for crude-oil distn. but which can be used for soap, naphtha, sulfonic acid, etc. The basic principle of the app. is that the liquid to be distd. is led in a uniform stream through the heating zone and from there is immediately brought to a dephlegmator filled with filter rings and to a condenser. Since the heating surface and the vaporization space are very large compared with the amt. of heated liquid, a continued heating of the liquid is avoided. M. McMARON

Laboratory oven for carrying out chemical reactions. E. BEHL AND E. WENIGER. *Chem. Fabrik* 1931, 194.—A Cu block oven is described in which about 93% of the inner space is held const. to $\pm 2^\circ$. J. H. MOORE

Liquid filters for the preparation of artificial sunlight. R. LUTHER. *Z. phys. Chem.* 134, 234-5(1930), *d. C. A.* 25, 1450. E. J. C.

Use of filter pencils. E. SCHWARTZ VON BERGMANN. *Mikrochem. Emich Festschr.* 1930, 268-74.—Filter pencils made of tightly rolled ashless filter paper placed in a glass tube serve for the sepn. of ppt. such as silica which have to be further treated after ignition. B. C. A.

A buret for titration with solutions affected by air. PETER DICKENS. *Chem. Fabrik* 1931, 185.—The buret is fused into the stopper of the stock bottle and is so connected that it is filled by pressure from a cylinder of H or N. J. H. MOORE

A sample extraction apparatus. ARTHUR G. MILLIGAN. *J. Soc. Chem. Ind.* 50, 144T(1931).—The app. used for the extrn. of rubber with $(CH_3)_2CO$, consisting of a siphon extn. tube suspended in the wide neck of the boiling flask, is improved by inserting between the flask and cooler a tube filled with small glass beads to heat the condensate to the b. p. before it drops into the extn. tube. J. H. MOORE

An apparatus for extracting liquids with solvents of higher specific gravity. P. H. PRACINSKY. *Chem. Fabrik* 1931, 193-4.—The usual form of the extn. vessel is modified by putting in the bottom a glass filter plate through which compressed air, or other gas, is blown to mix the liquid and solvent, thus hastening the extn. Some results and a few references are given. J. H. MOORE

Cooling apparatus. JUAN HALLER. *Chem. Ztg.* 55, 257-8(1931).—The app. is for quickly cooling and holding small quantities of liquid at any temp. down to -15° . It depends upon the evapn. of CS_2 by bubbling a current of air. If the liquid container is replaced by a coil, gas currents may be cooled. J. H. MOORE

An automatic balance. F. J. VEENSTER, C. H. HOFMANN AND C. V. GIVAN. *Science* 73, 424-6(1931).—A description is given of a device operating on the displacement principle in which the usual difficulties encountered in such app. are overcome, a sensitivity of 0.02 g at a load of 3 kg is given. If it is attached to a 200-g balance an accuracy of less than 0.005 g is possible. J. H. MOORE

Microbalance with proportional swings and damped oscillations. J. DONAT. *Mikrochem. Emich Festschr.* 1930, 39-42 *d. C. A.* 25, 1121.—The balance may be adjusted to give a movement of 1 division per 0.05 mg. and may be loaded up to 500 mg. B. C. A.

A device for producing a slow constant flow of liquid. R. J. DAVIES. *J. Sci. Instruments* 8, 110-1(1931)—The liquid is forced out of the container by displacement with Hg, the flow of which is regulated by passing it through a system of capillary tubes. A steady flow of 25-150 cc. per day may be obtained. J. I. MOORE

The construction of some new acid pumps. H. WINKELMANN. *Chem. Fabrik* 1931, 195-6—Two pumps are described, 1 based on the "2 liquid" principle in which the working parts are protected by an "oil cushion" on the acid, so arranged that there is no mixing of oil and acid. J. I. MOORE

Improved apparatus of Engler-Reusler. JAROSLAV FORMÁNEK. *Chem. Obsor* 6, 36-7(37 English)(1931)—The stopper with a small metallic drum is replaced by a simple cork with a larger bore for a small burner, it also contains 6 bores for the air. A small hollow glass burner, fitted on an Flenmeyer flask, is used. A moderate air current is blown into the inside tube of the burner, this makes the min. gasoline flame larger. By this arrangement 6-g gasoline can be burned in 1 hr. The small delivery tube is ground into the adsorption cylinder and drawn out at the mouth into an opening 2 mm in diam. The opening is provided on the sides with 4 bores approx. 1.5 mm. in diam., and thus a better distribution of the combustion products, collected in the adsorption vessel, is obtained. J. KUCERA

Spray driers. O. ZAHN. *Chem.-Ztg* 54, 973-5(1930)—Patents covering driers of this type are reviewed. The essential points in design are the atomization of the liquid and the course of the hot air current. In the author's "Ravo-Rapid" system the drying chamber is a vertical cylinder. The liquid is allowed to fall centrally onto a rapidly rotating disk, while the main air feed rises in an ascending spiral and thus passes transversely across the course of the droplets. Subsidiary air feeds assist the atomization, and the dried powder is collected at the bottom of the chamber. It is claimed to yield a very uniform product and to have a high thermal efficiency. B. C. A.

Ionization manometer for small pressures. RUDOLF SEWIG. *Z. tech. Physik* 12, 218-21(1931)—The app. is described in detail. Curves are given showing its adaptability to pressures from 0.01 to 0.0001 mm. Hg. WILLIAM F. VAUGHAN

The thermometer and its production. H. FISCHER. *Uhrmacherkunst* 1929, Nos. 7, 8, 9, *J. Soc. Glass Tech.* 14, 41-2A—The evolution of the thermometer is traced from Galileo's "thermoscope" (1593) to the first Hg filled thermometer of Fahrenheit (1714). Methods of making modern thermometers are described, the requirements for the glass and for the filling liquid are discussed, calibration is described, and constructional details of the various types of thermometer are given. E. H.

Theory, design and construction of sensitive vacuum thermopiles. C. H. CARTWRIGHT. *Rev. Sci. Instruments* 1, 592-604(1930). E. H.

Impervious tubes of pure alumina. F. ADCOCK AND D. TURNER. *J. Sci. Instruments* 7, 327(1930)—Tubes are made by extrusion through dies and fired at about 1900° in a C resistance furnace, being supported by packing with powdered alumina. They are then glazed by careful heating with an oxyhydrogen flame and are practically impervious to gases at the ordinary temp. B. C. A.

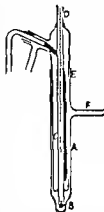
Sublimation apparatus using sintered glass. A. SOLTYS. *Mikrochem. Emich Festschr.* 1930, 275-9—An app. is described wherein saturation of the current of air by the vapor of the subliming substance is ensured by spreading the material on a plate of sintered glass, the max. sublimation velocity at any given temp. is thus obtained. The plate is fused into a glass tube, which is fitted at the lower end with a water-cooled tube for condensing the vapor and at the upper end with a capillary for the admission of the air current. B. C. A.

A strongly lighted fluorescence microscope. MAX HARTINGER. *Mikrochemie* [N. S.] 3, 220-3(1931)—In the search for a suitable source of ultra-violet rays, iron proved to be the most practical metal to use for the electrodes, improvements in the design were made whereby a steady light, free from flickering, could be obtained. The lamp finally designed uses about 5 amp. of current and by means of it good illumination and great detail is obtained under the microscope. The app. of improved design can be obtained from C. Reichert of Vienna. W. T. H.

New photoelectric recording microphotometer. J. A. CARROLL AND E. B. MOSS. *Monthly Not. Roy. Astron. Soc.* 91, 191(1930)—This is a detailed description with diagrams and photographs of a new recording microphotometer patterned after the Dobson-Skinner instrument and designed for studying photographs of stellar and solar spectrum lines and for measuring extended objects, as e.g., nebulae and the corona. J. W. McFARLANE

Fractional macro- and microsublimation at ordinary and at reduced pressures. G. ILLARI. *Ann. chim. applicata* 21, 127-36(1931)—Sublimation, although having

several advantages over distn and crstn, has not been developed to the same extent owing to the lack of suitable app. After defining the *initial temp of sublimation* as that temp at which the first sublimate is observed, for a given pressure, I. describes the app shown for detg this temp. A is a tube provided with a well B, in which the material is placed, and a side arm F, which may be connected to a suction pump. Inside is the condenser C, which passes through the ground glass joint at E and is provided with the thermometer D, which passes right through so as to be imbedded in the material at B. The material is heated at such a rate as to raise the temp 1° per min until the first sublimate is observed. In this way the *initial temp of sublimation* of the following were detd at 760 and at 7 mm: anthracene 98°, 50°, naphthalene 50°, 22°, benzoic acid 82°, 27°, cinnamic acid 95°, 42–4°, α -naphthol 67°, 26°, β -naphthol 78°, 27–8°, camphor 40°, 20°, caffeine 113°, 48°, theobromine 175°, 87–90°. Detns were also made at intermediate points for some of the above compds and the curves plotted. Results show that the rate of rise of the *initial temp* is most rapid between 0.7 to 100 mm, and then increases more slowly. By modifying the app, so that it is shaped like a U and provided with condensers with ground glass joints at the bases of the U, a mixt. of 2 compds. can be *fractionally* sublimed by opening only one of the condenser arms, till 1 component has sublimed and then opening the other to condense the second component.



A. W. CONTIERI

A simple thermostat for 20°, with temperature regulation independent of room temperature. V. ČUPR. 2 *Elektrochem* 37, 129–30 (1931).—There is a continuous flow of cool water (cooled with ice if necessary) through the thermostat, this eliminates the necessity for stirring. The water is heated before it enters the bath by an elec. heater which operates intermittently when the circuit is closed by the thermoregulator and relay system. An EtOH-Hg regulator is used, with a Hg relay. The regulation interval is said to be 0.02°.

CHANNING WILSON

Porous masses from gypsum [for filter stones, etc.] (Ger. pat. 522,139) 18.

Oil filters. SOC. ANON. DES ÉTABLISSEMENTS TICALÉMIT. Fr. 697,040, Sept. 27, 1929.

Filter presses. WM. H. YATES and JAMES A. BLACK. Fr. 697,314, June 13, 1930.

Endless sifting or filtering band. FRIED. KRUPP GRÜSONWERK A.-G. Ger. 522,165, Sept. 26, 1928.

Filter system for laundry apparatus. GEORGE W. DUNHAM (to Whirldry Corp.). U. S. 1,798,730, March 31. Structural features.

Apparatus for the continuous separation of liquids of different specific gravities. AKTIEBOLAGET SEPARATOR. Fr. 697,837, June 24, 1930.

Apparatus for determining the velocity of separation of small particles suspended in liquids. KARL T. R. LUNDGREN. Ger. 516,720, June 9, 1928.

Rectification system for separating constituents of gaseous mixtures such as those of air. CLAUDE C. VAN NUYS and JOSEPH L. SCHLATT (to Air Reduction Co.). U. S. 1,799,937, April 7. Various details of app. and operation are described. Cf. C. A. 24, 4963.

Device for removing dust from air by centrifugal action. CHARLES A. WINSLOW and ELBERT J. HALL. U. S. 1,798,510, March 31. Structural features.

Electromagnetic separator for separating magnetic from non-magnetic material. UKRAINSKOE AKTIONERNOE OBSHCHESTVO PO PATENTAM I REALIZATSI IZOBRETIENII UKAORIZ. Russ. 14,436, Mar. 31, 1930. An app. for cong. ores or separating copper shavings from iron and cast-iron shavings is described.

Modification of an electromagnetic separator. S. V. SAMBUROV. Russ. 19,186, Feb. 28, 1931. Addn. to Russ. 14,436 (preceding abstract).

Colorimeter. WM. L. PATTERSON (to Bausch & Lomb Optical Co.). U. S. 1,799,639, April 7.

Syringe hydrometers. LEO EDELMANN (to E. Edelmann & Co.). U. S. 1,800,138-9, April 7. Structural features.

Membrane decanting siphon. OSCAR PETERS. U. S. 1,799,023, March 31. Structural features.

Dehydrating agent suitable for use in desiccators. GEORGE F. SMITH. U. S.

- 1,798,175, March 31 A granular dehydrating material comprises $Mg(ClO_4)_2$ and $Ba(ClO_4)_2$.
- Drying apparatus. V. N. ANTONOV. Russ. 19,533, Feb. 28, 1931.
- Tubular drying apparatus. I. G. FARBENVIND A-G. Ger. 522,400, Mar. 7, 1928.
- Details of the air blast supply are described.
- Canal or chamber driers. MÖLLER & FREYER. Ger. 516,837, Mar. 4, 1928.
- Details of loading are given.
- Drying drum. EDUARD SEVTER. Ger. 522,140, Aug. 30, 1928. Details of the discharging means are described.
- Rotary drying drum for casein, etc. EDUARD SEVTER. Ger. 522,142 and 522,143, May 11, 1930.
- Oven for drying and burning sea weeds. N. A. ZEMLIANTZIN AND I. I. MIKHAILOVSKII. Russ. 19,533, Feb. 28, 1931. Mechanical details.
- Drier for beets, etc. I. G. ZLOBINSKII. Russ. 19,535, Feb. 28, 1931. Constructional details.
- Concentrating or evaporating apparatus. LÉON DEPASSE. Fr. 698,742, Oct. 14, 1929.
- Apparatus for continuous evaporation or distillation under vacuum. OTTORINO ANGELLO. Fr. 698,376, July 4, 1930.
- Evaporating milk or other liquids in multiple-effect evaporators with superposed flash chambers. DAVID D. PEEBLES. U. S. 1,799,478, April 7. Various details of app. and operation are described.
- Apparatus for the wet purifying and absorbing of gases. THOMAS BARTLETT HAYWARD Co. Ger. 516,990, Dec. 11, 1928. App. for spraying the gas with various liquids in various chambers is described.
- Hollow metal container for holding gases under high pressure. SYDNEY B. N. HILL (one-half to Chesterfield Tube Co., Ltd.). U. S. 1,798,545, March 31. Structural features.
- Apparatus for proportionate mixing of gases such as gas and air for burners. ROSSWELL W. THOMAS (to Phillips Petroleum Co.). U. S. 1,798,202, March 31. Various structural details are described.
- Gas washer with spray nozzles. FREDERICK H. WAGNER (to Bartlett Hayward Co.). U. S. 1,798,822, March 31. Structural features.
- Apparatus for washing smoke in a stack. JOHN BADAR. U. S. 1,799,084, March 31. Structural features.
- Tower for effecting contact between gases and solid materials as in purifying fuel gas. CHARLES COOPER (to W. C. Holmes & Co., Ltd.). U. S. 1,798,507, March 31.
- Tower for effecting contact between liquids and gases as in washing fuel gases. SVEN G. STYRUD (to Koppers Co.). U. S. 1,799,563, March 31.
- Device for mixing fluids such as steam and water in streams flowing together past a baffle. ELVINO O. NELSON (one-third each to Leslie R. Johnson and Charles N. Morrison). U. S. 1,798,336, March 31. Structural features.
- Modification of a cylindrical container for drawing samples of liquids. S. V. BELITSKII. Russ. 19,373, Feb. 28, 1931. Addn. to Russ. 18355. Structural details.
- Emulsifying device suitable for use with pipe or hose lines. FRED E. LEBARON. U. S. 1,799,051, March 31. Structural features.
- Apparatus for charging water or other liquids with carbon dioxide. MAX A. GUGGENBUHL. U. S. 1,799,295, April 7. Structural features.
- Apparatus for carbonating solutions (such as alkaline brines for obtaining sodium bicarbonate). HEINZ H. CHIESNA. U. S. 1,799,354, April 7. Structural features.
- Pressure regulator. ROSE E. JOYCE. U. S. 1,798,377, March 31. Structural details.
- High-speed beating mill. ERWIN KRAMER (to the firm Hartmann-Metall A-G). U. S. 1,798,886, March 31. Structural features.
- Apparatus for treating materials with hydrogen under high temperature and pressure. DEUTSCHE BERGHEIM A-G FÜR KÖHLE- UND ERDÖL-CHEMIE (Wolfgang Grote and Julius Scherenbeck, inventors). Ger. 516,993, Nov. 8, 1927. The high-pressure vessel and tubes are made of C Fe coated with Be.
- Apparatus for liquefying chlorine. KRAMER & Co. G. M. & H. Fr. 698,443, July 5, 1930.
- Apparatus for the decomposition of water. LOUIS FOMÉ and MICHEL DUMAS. Fr. 698,113, June 27, 1930. An app. is described consisting of 2 vessels one inside the other, the inner vessel containing catalytic materials composed of thin layers of Al alloyed with Sn or Zn and $FeCO_3$ combined with CaH_2 for splitting water into H and O.

Digesting or extracting plant. JENS W A ELLING Fr 697,406, June 6, 1930. A plant for digesting extg or washing materials is described in which the material is exposed to heat under alternatives of increase and decrease of pressure

Apparatus for washing and flocculating slimes. ANTOINE FRANCE Fr 698,463, July 7, 1930

Apparatus for measuring the opacity of translucent sheets. BROOKS I. HILL. U S 1,798,203, March 31 Various mech, optical and elec features are described

Apparatus for casting various articles from plastic or liquid masses. JENS J. JENSEN (to Boggild & Jacobsen) U S 1,798,326, March 31 Mech features

Photoelectric cells. SOC. DES LAMPES PHOS (Georges L C L Déjardin and Maurice P Warin, inventors) Fr 697,330, June 13, 1930 The cathode of a photoelec cell is composed of a heavy deposit of K covered with a thin layer of Cs It may be sensitized by a luminous discharge in H

Photoelectric cells. SOC. DES LAMPES PHOS (Georges L C L Déjardin, inventor) Fr 697,331 June 13 1930 In a photoelec cell the K is deposited in a very thin layer on a layer of superficially oxidized Mg

Photoelectric cells. VEREINIGTE GLÜHLAMPEN UND ELECTRICITÄTS A G Austrian 121,656, Oct 15, 1930 Manipulative details of the manuf are described

X-ray apparatus for examining crystalline substances. N V PHILIPS' GLOEI-LAMPENFABRIEKEN Fr 697,731, June 20, 1930

Ultra-violet lamp. PALMER H CRAIG U S 1,798,658, March 31

Producing vacuums in devices such as radio tubes or valves. HUGH S COOPER (to Kemet Laboratories Co) U S 1,800,131 April 7 For 'cleaning up' gases, there is introduced into the envelope a pellet of powder alloy comprising Mg with 5-30% of an alk earth metal, followed by evacuating and flashing

Acetylene generator. MESSER & CO G M B H. Ger 516,907, Dec. 3, 1929. The app is designed for CaC_2 dust, which falls from a hopper onto a revolving plate, which sprays it into a second hopper whence it is discharged into mechanically stirred water

Acetylene generator. AUTOGENWERK SIRIUS G M B H Ger 520,552, Feb 6, 1930 Addn to 474,145 (C A 23, 2855)

High-pressure acetylene generator. LUDWIG ADLER G M B H and RICHARD CLEMENS. Ger. 522,492, Jan 16, 1930

Acetylene generator (water-to-carbide type.) WILHELM HAAR & CO Austrian 122,019, Nov. 15, 1930.

Thermostatic valve-control device. GASTON A BRUNELLE and CAMILLE BRUNEL. U. S 1,799,407, April 7. Structural features.

Thermostatic device for control of electric circuits. LEONARD SATCHELWELL. U. S 1,798,854, March 31 Structural features

Thermostatic device suitable for use in rectifying lubricating oil. WALTER B CLIFFORD (to Clifford Mfg Co) U S 1,798,192, March 31 Structural features.

Thermostatic flue damper regulator. ARTHUR C. McWILLIAMS U. S 1,798,431, March 31. Structural features

Circular five-chamber furnace. A. S FILIPOV Russ 19,517, Feb 28, 1931. Constructional details

Rotating furnace. RENÉ P É. ROMAN and THÉODORE LAFITTE Fr. 697,373, April 23, 1930 Improved construction is described

Shaft furnaces. SOC. ANON LE NICKEL Fr. 697,962, Sept. 30, 1929 Means is described for heating the air by heat from the combustion products

Furnace with an adjustable inner fire wall. ALFREDO NENDZA Ger. 516,904, Mar 16, 1928

Tunnel furnaces. WOODALL DUCKHAM (1929), LTD, and ARTHUR McD DUCKHAM Fr 697,639, June 18, 1930

Tunnel muffle kiln. LUDWIG RIEDHAMMER. Ger. 516,858, Jan 20, 1929. The firing box is suspended

Grate furnaces. STOCKHOLMS AKTIEBOLAGET PRIVAT (Iwan Arbatsky, inventor) Fr 697,688-89-90-91, June 19 1930 Means and app for lighting and working the furnaces are described

Inclined grate furnace. ERNST VÖLCKER Ger 516,699, Nov. 4, 1926

A charging and drying shaft for inclined grate furnaces. ERNST VÖLCKER. Ger. 516,849, Nov 7, 1926. The fuel is dried and heated by flue gas Cf C A. 24, 3932.

Furnaces using powdered coal. ALLGEMEINE ELECTRICITÄTS-GES. Fr. 697,836, June 24, 1930

Burner for powdered fuel. JAY G COUTANT Fr 698,697, Oct 8, 1929

- Regenerative gas furnaces. HERMANN MOLL. *Fr.* 697,743, June 21, 1930
 Low-temperature gas burner. FRANKLIN C. CARTER. U. S. 1,798,785, March 31.
 Burner for liquid fuel. MAX COBLITZ. *Fr.* 698,077, June 26, 1930
 Rotary-plate device for supplying fuel to furnaces, etc. KARL BESTA. *Ger.* 522,458, July 13, 1929
 Method of rapid charging of furnaces. HENRI GARREAU. *Fr.* 698,092, Oct. 7, 1929.
 Apparatus for leading heated air to furnaces by tubes located over the fire chamber. SVENGE G. SYVERSEN. *Ger.* 516,905, Feb. 5, 1928
 Rotary annular tray furnace for continuous heat treatment of pulverulent materials such as fuel subjected to low-temperature distillation. FRIEDRICH BARTLING. U. S. 1,798,649, March 31. Structural features
 Heating apparatus suitable for retorts, leers, etc. FRANZ PUENING. U. S. 1,799,702, April 7. Structural features of an app. in which flow of hot gases is periodically reversed
 Combined air, water and flue-gas heat interchanger. ALBERT E. LECK. U. S. 1,798,330, March 31. Structural features.
 Heat-exchange apparatus suitable for use with gases. ANTHONY CONEJOS. U. S. 1,799,039, March 31. Structural features.
 Heat-exchange apparatus suitable for use with liquids. PERCY C. KEITH, JR. (to Refinery Engineers, Inc.). U. S. 1,799,620, April 7. Structural features.
 Heat-exchange apparatus suitable for use as a reboiling element for fractionating towers. WALTER M. CROSS (to Gasoline Products Co.). U. S. 1,799,734, April 7. Structural features.

2—GENERAL AND PHYSICAL CHEMISTRY

FREDERICK L. BROWNE

- Henry Gall. LÉON GUILLET. *Rev. metal.* 23, 183-4(1931).—Biography with portrait. E. H.
 Emil Warburg's 85th birthday. EDGAR MEYER. *Naturwissenschaften* 19, 241 (1931). B. J. C. VAN DER HORVEN
 M. W. Beijerinck, 1851-1931. A. J. KLUJVER. *Nederland. Tijdschr. Hyg. Microbiol. en Serol.* 4, 173-81(1931).—An obituary with portrait. J. C. JOEKES
 Anne Wilhelm van der Haar. D. H. WESTER. *Chem. Weekblad* 28, 272-5 (1931).—Obituary with portrait and list of publications. G. G.
 Dr. Willy Herzberg. PAUL RÖDERS. *Chem.-Ztg.* 55, 317(1931).—Obituary. E. H.
 Jacobus Henricus van't Hoff. Memorial address on the twentieth anniversary of his death. E. ABEL. *Oesterr. Chem.-Ztg.* 34, 54-6(1931). Personal recollections of van't Hoff. D. BADER. *Ibid.* 55-9. E. H.
 Josef Schneider, 1864-1931. JAROSLAV SCHNEIDER. *Chem. Obsor.* 6, 33-6(1931).—Obituary with a portrait. JAROSLAV KUČERA
 F. W. Semmler. H. BRUCKER ROSE. *Z. angew. Chem.* 44, 301-2(1931).—Obituary. E. H.
 Reform of the chemical nomenclature. XI. G. SIMONI. *Boll. chim. farm.* 69, 891-4, 897-900(1930) XII. *Ibid.* 70, 49-55(1931) XIII. *Ibid.* 89-90 XIV. *Ibid.* 139-40, cf. *C. A.* 25, 240. G. SCHWOCH
 Should chemistry be taught by the deductive method followed by the empirical-
 deductive method? W. H. VAN MELS. *Chem. Weekblad* 28, 260-2(1931). G. G.
 The achievements of Josef Jan Fré in polarimetry. K. V. ZIELECKI. *Listy Cukrovar.* 49, 333-4(1931).—A review of the contributions of Fré to the construction of polarimeters. FRANK MARSH
 The distribution of ozone in the atmosphere. D. CHALONGE and E. DUBOIS. *Compt. rend.* 192, 808-10(1931).—It was previously concluded that the O₃ of the atmosphere is concentrated in a relatively thin layer at a high altitude. The present photographic and theoretical study of the solar spectrum, however, indicates that the distribution of O₃ in the atmosphere is more complex, and that appreciable quantities of it exist at relatively low altitudes. R. H. LOMBARD
 Thermal energy of water in the Arctic regions. H. BARJOT. *Trans. 2nd World Power Conference, Berlin* 16, 3-15(1930). E. J. C.
 The experimental basis of the international temperature scale in the low-temperature range. W. H. KESSOM. *Z. ges. Kälte-Ind.* 38, 20-1(1931); cf. *C. A.* 24, 2341-2—

reply to Henning (*C. A.* 24, 5592). The latter's report causes no change in the standard of the Leyden Inst., which considers the international temp. scale to be reproducible to $\pm 0.02^\circ$. A comparison of more thermometers of various makes is necessary, and such an international comparison is in process. Until the results of this work become known, the Leyden Inst. will continue to employ the approved method of making a direct calibration of the Pt thermometer used with the gas thermometer.

F. D. ROSSINI

A physicochemical study of some organic compounds solid at ordinary temperatures, proposed as temperature standards. F. BURDET. *Annales soc. esp. fis. quim.* 29, 89-125 (1931). —For the prepn. of the International Pharmacopeia a set of thermometric standards melting or boiling from ordinary temp. to 350° was desired. The substances chosen were fractionally crystd., the d was detd. by Jena-glass dilatometer and viscosity and superficial tension were detd. first by the Poiseuille method, then by capillary rise, both with the same app. Heat of combustion was detd. by the Bureau des 1 talons method. The substances chosen were salol m 41.68° , benzophenone m 47.85° , b 305.9° , naphthalene m 80.06° , b 217.96° , benzoic acid m 122.45° , b 250.0° , phthalic anhydride m 131.60° , b 245.1° , mannitol m 166.00° , anthracene m 216.0° , b 329.9° , carbazole m 240.3° , anthraquinone m 241.8° , b 376.8° .

I. M. SYMES

Application of a squared diagram to the representation and calculation of equilibrium in a gas-liquid reaction. PIERRE MONTAGNE. *Compt. rend.* 192, 677-9 (1931). —In reactions of the type, $CO + H_2O = CO_2 + H_2$, the use of a squared diagram permits rapid calcn. of equl. concns. at different temps.

DOV BAUTSP

The polishing of surfaces. J. M. MACAULAY. *J. Roy. Tech. Coll. (Glasgow)* 2, 379-85 (1931). —Various theories regarding the nature of surface polish are discussed. Exptl. data are shown bearing on the theory that in the process of polishing the surface material is actually raised to its melting temp., giving the liquid like finish to the polished surface. When Pb_3O_4 , PbO_2 , and $CaCO_3$ are used as polishing powders evidence was found that they become changed to lower oxides, PbO , CaO , etc. This is considered as more or less direct evidence of high temp. during the polishing process.

A. L. K.

Color measurement and spectrometric color determinations. J. A. A. KETELAAR. *Chem. Weekblad* 28, 132-3 (1931).

F. J. G. DE LEEUW

Color measurement and color standards. M. J. SCHORP. *Chem. Weekblad* 28, 133 (1931), cf. *C. A.* 25, 1670.

F. J. G. DE LEEUW

From the atomic weights of Dalton to the isotopes of Soddy. L. DE LOSVAY. *Természettudományi Közlemények* 63, 193-200 (1931). —Historical address.

S. S. DE F.

Development and present state of our knowledge of the natural system of the elements. F. PANETH. *Chem. Rundschau Mitteleuropa u. Balkan* 7, No. 15, 98-100 (1930). —A general review.

S. S. DE F.

The atomic weights of nitrogen and silver. I. The ratio of ammonia to silver. GREGORY P. BAXTER AND CHARLES H. GREENE. *J. Am. Chem. Soc.* 53, 604-13 (1931). —The ratio Ag/NH_3 was detd. as 623420 by weighing NH_3 adsorbed on dehydrated chabazite in a closed tube, causing a suitable quantity of NH_3 to evaporate into dil. HCl or HBr and measuring the halogen content nephelometrically. By using this ratio and that for NO_3/Ag the at. wt. of N was calcd. as 14.0078 and that of Ag 107.879.

L. P. HALL

Molecular and atomic volumes. XXIV. Technical experiences in volume measurements of density at low temperatures. E. WÖNNENBERG, WERNER FISCHER, ADOLF SÄPPER AND WILHELM BILTZ. *Z. physik. Chem., Abt. A*, 151, 1-12 (1930); cf. *C. A.* 24, 2343, 4200. —Practical matters concerning the app. temp., filling gas and handling of the materials previously investigated are discussed. The corrections for compression of the filling gas are given in detail, as well as the absorption of the filling gas by various substances. XXV. The volume occupied by crystalline organic compounds at low temperatures. WILHELM BILTZ, WERNER FISCHER AND E. WÖNNENBERG. *Ibid.* 13-55. —The vols. of various typical cryst. org. compds., contg. H, or both O and H, were detd. at temps. down to -195° . The results are tabulated according to the principles of additivity, effect of constitution (arrangement of atom and at. groups), and equal vols. XXVI. The molecular volumes of several silver and potassium salts of monobasic aliphatic acids. W. FISCHER AND ALFRED LEMKE. *Ibid.* 56-64. —The densities of 9Ag and 3K salts of normal monobasic aliphatic acids were detd. at room temp. and at -78° . The measurements were carried out pycnometrically, with toluene as the enclosing liquid. The mol. vols. of the salts are discussed in the light of Biltz's space theory. XXVII. The space requirements of several cellulose derivatives and the gas absorption of acetylcellulose. E. WÖNNENBERG, W. FISCHER AND W. BILTZ. *Ibid.* 65-70. —The d and coeffs. of expansion of cellulose, cellulose acetate and cellobiose were measured. Some

data of gas absorption (air and H_2) by cellulose acetate are given. Air is absorbed about 3 times as much as H_2 , obeying Henry's law XXVIII. The molecular volumes of several substituted benzoic acids. LI KLEMM AND WILHELM KLEMM *Ibid* 71-9 — The mol vols. of methyl, hydroxy, amino-, nitro-, chloro- and bromo-benzoic acids were detd

LOUIS WALDBAUM

Electron diffraction and molecular structure. R. WIERL *Ann Physk* 8, 521-64(1931) — It is shown that if one gives to a fast moving electron a wave length according to de Broglie, then the theoretical treatment of Debye on the x ray diffraction by gases may be applied to electron diffraction by gases. The diffraction pattern obtained on allowing the electron beam to penetrate the vapor of a gas which is forced from a narrow nozzle into the electron beam at right angles is characteristic of the structure of a single mol of the vapor under investigation. By allowing the time of exposure of the photographic film to be only a fraction of a sec. the theoretical treatment becomes sufficiently accurate to det. the form and the at. distances of the mol. and thereby gives to the structures of stereochemistry the clearest proof of their spacial reality. The results are as follows

Molecule	At. distances A U	At. form
Br_2	Br—Br 2.28	
CO_2	O—O 2.26	linear
CS_2	S—S 3.16	linear
SO_2	S—O 1.37	angular?
N_2O	N—N 2.38	
CCl_4	Cl—Cl 2.98	tetrahedral
$GeCl_4$	Cl—Cl 3.43	"
$SiCl_4$	Cl—Cl 3.29	"
$TiCl_4$	Cl—Cl 3.61	"
$SnCl_4$	Cl—Cl 3.81	"
$CHCl_3$	Cl—Cl 3.04	
CH_2Cl_2	Cl—Cl 3.16	Angular, C—Cl angle about 110°
CBr_4	Br—Br 3.35	
BCl_3	Cl—Cl 3.03	
PCl_5	$\left\{ \begin{array}{l} Cl—Cl \\ P—Cl \end{array} \right. \begin{array}{l} 3.18 \\ 2.04 \end{array}$	pyramidal
C_6H_6	C—C 1.39	all C atoms in a plane
C_6H_{12} (cyclohexane)	C—C 1.51	angular
C_5H_{10} (cyclopentane)	C—C 1.52	a plane
C_4H_8		
C_3H_6	C—C 1.5	tetrahedral angles
C_2H_4		

MALCOLM DOLE

The effective cross section of argon and hydrogen against electrons of 0.2 to 6 volts. II. GAERTNER *Ann Physk* 8, 135-62(1931) — The effective cross section of A was accurately detd. for various electron velocities. The min cross section occurs at 0.4 v instead of 0.7 v as previously reported by Rusch (*C. A.* 20, 3129). The effective cross section in sq. cm. per cc. of gas at a pressure of one mm. of Hg at 0° decreases from 24 at 5.5 v to about 1 at 0.4 v and then rises again to 6 at 0.17 v. The effective cross section of H_2 rises slowly from 32 at 6 v to 39.5 at 2.5 v and then decreases to 30 at 0.3 v.

P. T. NEWSOME

Refractivity of a binary mixture and its relation to the molecular size of the components. TETSUYA ISHIKAWA *Anniversary Vol. Dedicated to Masumi Chikashige* (Kyoto Imp. Univ.) 1930, 275-91 — Formulas relating n and the compn. of a binary mixt. are reviewed and discussed. For successive members of a homologous series the equation,

$$\sigma_2 = \sigma_1 \sqrt{\frac{\beta_1 \sigma_2}{M_1 \rho_1}}, \text{ is proposed connecting the mol diam., mol wt and sp vol. The mol}$$

diam. of one component of a binary mixt. can be calcd. from the refractivity provided that of the other is known. The equation is strictly true for MeOH, EtOH and n - C_4H_9 -OH nearly so for iso- C_4H_9 -OH and iso- C_5H_{11} -OH, but not true for n - C_5H_{11} -OH and n - and iso- C_6H_{13} -OH. This suggests that the addn. of the CH_2 -group to the alc. series above C_4H_9 -OH is not literally an elongation of the C chain. Expts. on the refractivity of the alcs. with benzene and of the alc. pairs are described. The values of σ obtained are compared with those from the expression of Lorentz-Lorentz. ALLEN S. SMITH

X-ray studies of motions of molecules in dielectrics under electric stress. RALPH

D. BENNETT *J. Franklin Inst.* 211, 481-7(1931)—X ray examn of paraffin solidified under elec. stress indicates that the thin flakes of the hydrocarbon crystal with long chains perpendicular to the face of greatest area orient themselves edgewise to the applied elec field (So that the elec field is perpendicular to the plane of the hydrocarbon chain)

MALCOLM DOLE

A general theory of paramagnetic rotation in crystals. H. A. KRAMERS. *Proc Acad Sci Amsterdam* 33, 959-72(1930), cf *C. A.* 25, 873—Math analysis shows that in a doubly refracting crystal of slight anisotropy possessing magnetic rotatory power the introduction of a vector of rotation permits the simple description of properties of a luminous beam of any orientation. A consideration of the properties of an atomic system with an odd number of electrons in an external field of purely elec. origin shows the energy levels are necessarily doubly degraded. The magnetization produced in an atom possessing such a degraded field by a magnetic field of arbitrary direction leads to a paramagnetic rotation in a crystal contg. such atoms.

A. P. SACHS

Electrical properties of molecules. II. A. H. W. ATEN. *Chem Weekblad* 27, 235-6(1930); cf *C. A.* 24, 3143—Dipole moments of org. compds. are discussed.

H. S. V. KLOOSTER

Magnetic rotatory power of halogen derivatives of saturated hydrocarbons in the gaseous state. R. DE MALLEMAN AND P. GABLANO. *Compt. rend.* 192, 278-80(1931), cf *C. A.* 25, 1418—The magnetic rotation in the gaseous and liquid states for the 578 mμ line was detd. for MeCl, EtCl, MeBr, EtBr, PrCl, iso-PrCl, CHCl₃, CCl₄ and PrBr. The mean values of the at. rotatory powers of H, C, Cl, Br and I are 5.5, 10, 33, 60 and 143×10^{-2} , resp., whereas the values ions are 108, 208 and 450×10^{-2} for Cl⁻, Br⁻ and I⁻. The rotatory power of the ion therefore is greater than that of the corresponding atom. The ratio of the ionic to the at. rotatory power is 3.2.

F. W. LAIRD

The dielectric constant of hydrogen chloride from 85° to 160°K. RICHARD M. COLE, GEORGE H. DENISON AND JACOB D. KEMP. *J. Am. Chem. Soc.* 53, 1278-82(1931)—The dielec. const. for solid HCl changes isothermally from 3 to 10 at the transition point at 98.4°K. This behavior is predicted by Pauling's theory of the free rotation of mol. in crystals.

JOHN R. HILL

The dielectric properties of antimony pentachloride and phosphorus pentachloride. J. H. SIMONS AND GILBERT JESSOP. *J. Am. Chem. Soc.* 53, 1263-6(1931)—This work was done with a view to obtaining information regarding the mol. structure of PCl₅ and SbCl₅. The molar polarizations of SbCl₅ and PCl₅ were measured in CCl₄ solns. and the dipole moments concluded to be either zero or very small. The dielec. consts. of SbCl₅ and PCl₅ were measured in both liquid and cryst. states. The specific cond. of PCl₅ was measured. Various structural theories for PCl₅ and SbCl₅ previously advanced by other authors are discussed. The authors conclude from their own data that these compds. have a symmetrical structure with a ten-electron shell for the central atom.

R. H. C.

Magnetism and molecular structure. I. The magnetic susceptibilities of some inorganic sulfides and electronic isomers. S. S. BHATNAGAR AND S. L. BHATIA. *Proc. 15th Indian Sci. Congr.* 1928, 141, cf *C. A.* 23, 4857, 24, 5550—The magnetic const. were detd. (at room temp.) of the yellow and red sulfides of As (solids) and also of their colloidal solns. in water. A modified Wilson's magnetic balance was used. The results (not given) show that realgar (As₂S₃) is less diamagnetic than orpiment (As₄S₃) both in the solid form and in colloidal solns. These differences can be explained on the basis of atomic structure and the valency theory of Langmuir.

E. J. C.

Electric moment and molecular structure. III. Double and triple bonds and polarity in aromatic hydrocarbons. C. F. SMYTH AND R. W. DUKSTY. *J. Am. Chem. Soc.* 53, 1296-304(1931), cf *C. A.* 25, 2032—The dielec. consts. and ds. of dil. solns. of the Pb substituted ethylenes and acetylenes were measured between 10° and 70°. The mol. refractions were also detd. and the data used to calc. the elec. moments of the mols. The Pb substituted ethylenes have small or zero moments according to their symmetry about the double bond. The replacement of H on a doubly-bonded C by a hydrocarbon group gave moments of the same magnitude as those resulting from a similar replacement in C₆H₆. The large moment for PhC≡CH and still larger value for PhC≡CPh give evidence of unsymmetrical mols., this is explained in terms of a tautomeric equil. between a symmetrical form and an unsymmetrical contg. bivalent C, the shift in equil. with temp. causing the temp. variation of the moment observed.

H. W. L.

The magnetic susceptibilities of the polyoxymethylenes and formaldehyde solutions. W. GOOD. *J. Roy. Tech. Coll. (Glasgow)* 2, 401-9(1931)—The investigations described were undertaken in the hope that measurement of the magnetic susceptibilities of the polyoxymethylenes and of formaldehyde solns. would furnish some information

regarding the constitution of highly polymerized substances. The values found for the polyoxymethylenes are not in agreement with the theoretical values calculated from the formulas that have been proposed to represent their constitution. They seem to lend support to the view that the polymers consist of small elementary units held together by strong crystal lattice forces. The values found for formaldehyde solns. support the view that the hydrates HOCH_2OH and $\text{HOCH}_2\text{OCH}_2\text{OH}$ are in equal in formaldehyde solns. A. L. KIMLER

Avogadro's number and "mean free path." SATYENDRA RAY. *Proc. 15th Indian Sci. Congr.* 1928, 141.—Millikan calls the mean free path a "hypothetical" quantity. It is shown that this mean free path is not related in any intimate manner with the properties of a gas like vol. pressure temp. or entropy, if the expression usually accepted for it $\lambda = 1/\sqrt{2}n\sigma$ is correct. Instead the fairly close agreement of Avogadro's no. for different atoms means that the size of atoms is very nearly identical and that perhaps $\lambda = \sqrt{3} \cdot a$ is a more nearly correct expression for the mean free path than the one usually accepted which makes the value of λ infinite when the size of particle is infinitely small. L. J. C.

Viscosity, heat conduction and diffusion in gas mixtures. XIII. Diffusion constants of dilute gas mixtures. MAX TRAUTZ AND WILHELM RIES. *Ann. Physik* 8, 193, 54 (1931) cf. C. A. 25, 2034.—The diffusion constants of C_2H_2 -H₂ and CCl_4 -H₂ mixts. were detd. at 23°. The diffusion const. (D) was unaffected by changing either the size of the tube or the rate of flow. The following values were obtained: $D_{\text{C}_2\text{H}_2\text{-H}_2} = 0.363$ sq. cm. sec. $D_{\text{CCl}_4\text{-H}_2} = 0.342$ sq. cm. sec. P. T. NEWCOMB

Methods for measuring the mobilities of gaseous ions. PETER GOLDMARK AND FRANZ KAMMEL. *Sitzb. Akad. Wiss. Wien, Abt. IIa* 139, 241-53 (1930).—By an improvement in the Rutherford Frank method, the mobility of gaseous ions can be measured without extrapolation. Math. proof and a description of the app. are given. Sp. ionic mobilities (μ) for atm. ions were detd. (1) by means of a sine form alternating voltage for the pos. ion, $\mu = 1.364 = 0.009$ sq. cm./v. sec., and for the neg. ion, $\mu = 1.811 = 0.008$ sq. cm./v. sec., (2) by means of an abruptly changing alternating voltage for the pos. ion $\mu = 1.379 = 0.008$ sq. cm./v. sec., and for the neg. ion, $\mu = 1.814 = 0.008$ sq. cm./v. sec. DON BROUWER

Uniform propagation of flame. N. R. SEN AND H. K. SEN. *Nature* 127, 125-6 (1931).—The velocity equation shows the condition for uniform propagation to be $\theta = \frac{\alpha Q}{\rho C}$ where θ = the ignition temp., ρ = density of the inflammable mixt., C = the sp. heat and αQ = the fraction of the total heat of combustion utilized for conduction. The above equation takes note of the distribution of temp. in every section of the gaseous medium comprising the traveling flame. The traveling flame is also considered as a region having temp. gradients corresponding with the combustion temp. at the rear and the ignition temp. at the front. In advance of this a progressive fall of temp. to that of the cold gas occurs. The ratios of flame to ignition temps. were for: Coal gas + air 2.62, CH_4 + air 1.91, C_2H_6 + air 1.90, CO + air 2.08. FRANK MARENH

Highly dilute flames of alkali metal vapors with hydrogen halides. GÉZA SCHAY. *Z. physik. Chem., Abt. B* 11, 291-315 (1930), cf. C. A. 29, 3221, 20, 1947, 22, 1898.—The vapors of Na and K under low pressure react with HCl, HBr or HI with displacement of H atoms. These atoms react with H₂ to form H₂ and Na or K to yield Na + H₂ with the emission of light. The velocity of the reaction of sodium with hydrogen halides. H. V. HARTEL. *Ibid.* 316-20 (1930).—The reaction of Na vapor with HI, HBr and HCl was followed by a diffusion method (C. A. 25, 1726). At 240° Na and HI react at every collision, Na and HBr after 6 collisions, Na and HCl after 86. The heats of activation for their reactions were calcd. as 200, 1900 and 4500 lg.-cal. and reaction velocities were detd. as 53,000, 10,000 and 800 mols./sec. $\times 10^{-12}$, resp. A similar heat of activation of Na and HCl was also found from measurements at different temps. L. P. H.

Highly dilute flames. GÉZA SCHAY. *Fortschr. Chem. Physik physik. Chem.* 21, 68 pp.—A general review of the technique and interpretation of the reaction of dil. vapors of alkali metals with halogens. H. halides, Hg halides and stannic halides. A bibliography is given. L. P. H.

The effect of an electric field on flames and their propagation. BERNARD LEWIS. *J. Am. Chem. Soc.* 53, 1304-13 (1931).—The effects of an elec. field on the flames of 10 hydrocarbon air and CO-air mixts. were investigated. The field was applied between fine wire gauze electrodes placed on either side of a stationary flame confined in a Pyrex tube. The flame is invariably pulled toward that electrode on which the charge is neg., thus moving in the direction of the pos. ion flow, its propagation is slowed down or

speeded up depending upon the direction of the elec. field. With appropriate direction and strength of field the flames of all the mixts. could be extinguished. P. H. E.

Calculation of the specific heats of gases from vapor-pressure curves. M. TRAUTZ AND W. BADSTÜBER. *Ann. Physik* 8, 185-202 (1931).—Sp. heats of paraffins, alcs., fatty acids, esters, ketones, aromatic bases, org. and inorg. halogen compds., etc., were calcd. from the relation $C_p = C_v + (\Delta L / \Delta T)$. Known values of C_p were used, and values of $\Delta L / \Delta T$ were detd. by application of the Clausius equation to the vapor pressure curves. Good agreement was obtained between calcd. and observed values of C_p for 30 different substances. C_p values were also calcd. for 43 other substances for which the C_p values have not been directly detd. The relation $C_p = C_v + 12$ gives good results for $C_p > 20$ and becomes more accurate as C_p increases. P. T. NEWSOME

The indices of refraction of liquids. MORTON MASTIS AND W. E. LAWTON. *J. Optical Soc. Am.* 21, 252-9 (1931). cf. *J. Optical Soc. Am.* 20, 271 (1930).—Math. considerations of the use of the av. of the 3 min. deviations and the use of the interpolation formula are given together with equations from which the n of a liquid enclosed in a hollow prism or of different liquids in the same prism can be calcd. The covers are not necessarily true plates, but the 3 angles of the prism should not differ by more than $10'$ from 60° . H. W. WALKER

The influence of elastic waves of thermal agitation on the interior pressure of liquids. TOKIO TAKEUCHI. *Proc. Phys. Mats. Soc., Japan* 13, 17-8 (1931).—A math. derivation of a relation between the pressure term in the van der Waals equation of state and the interior pressure of liquids. P. H. EMMETT

The volume of eighteen liquids as a function of pressure and temperature. P. W. BRIDGMAN. *Proc. Am. Acad. Arts and Sciences* 66, 180-200 (1931).—A new method of measuring compressibility and thermal expansions of liquids is described in which the liquid is enclosed in a "siphon" (a flexible metal bellows) which is then exposed to external hydrostatic pressure and the vol. change detd. from the change of length of the siphon. This method is applied to 18 liquids at 0°, 30° and 95° up to a pressure of 12,000 kg. per sq. cm., or to the freezing pressure and the results are collected in extensive tables giving the vol. as a function of pressure and temp. over this range. Compressions as great as 30% were measured. The pressure coeff. $(\partial p / \partial V)_T$ is not a function of vol. only and suggestions are made as to the theoretical significance of this. The liquids measured are pentane, isopentane, hexane, 2-methylpentane, 3-methylpentane, 2,2-dimethylbutane, 2,3-dimethylbutane, heptane, octane, decane, benzene, chlorobenzene, bromobenzene, carbon tetrachloride, bromoform, isopropyl alc., butyl alcohol and hexyl alcohol. Check measurements with previous methods were made with ether, water and water-glycerol mixt. (1:1). A. L. KILBER

The mixing contraction of liquids. NIKOLAI GERASIMOV. *Physik. Z.* 32, 226-9 (1931).—A math. discussion. C. E. P. JEFFREYS

Visible volume change in castor oil and Baku lubricating oil at low temperature and high pressure, studied by increase in viscosity and thus limited by clogging narrow tubes. G. TAMMANN AND A. PAPE. *Z. anorg. allgem. Chem.* 197, 90-2 (1931).—An anomalous pressure-vol. change is observed for castor oil at a temp. of -21° and 3000 kg./cm.² which is not found above 0°. The effect is not due to crystn. but is traced to a stopper effect in the small capillary leading to the compression cylinder. The same effect and cause were noted for Baku lubricating oil. R. H. LAMBERT

Boiling-point nomograph. D. S. DAVIS. *Chemist Analyst* 20, No. 3, 7-8 (1931).—A nomograph is shown, based on Trouton's rule applied to the Clausius-Clapeyron equation, which enables one to find at once the b. p. at atm. pressure when given at some other pressure, or vice versa. W. T. H.

Remarks concerning the work of W. Herz on boiling points and vapor-pressure formulas of organic liquids. E. W. MADGE. *Z. anorg. allgem. Chem.* 195, 335 (1931); cf. *C. A.* 24, 554.—It was shown by Herz that the vapor pressure of org. liquids is given by the formula $\log p = -(A/T) + C$ and that for 2 different liquids $(A_1/A_2) = (T_{b1}/T_{b2})$ where T_{b1} and T_{b2} are the abs. b. ps. at one atm. pressure. Madge points out that the latter relation follows from Trouton's law since A is proportional to the heat of vaporization. P. T. NEWSOME

Influence of the boiling temperature on the composition of azeotropic mixtures. B. KAMIEŃSKI. *Russk. Khim.* 11, 1-11 (in German 12) (1931).—Merriman's qual. rule (*C. A.* 8, 552) concerning changes in compn. of an azeotropic mixt. with change of the boiling temp. does not always hold. Vreva's considerations (*C. A.* 7, 3060) are too complicated. K. replaces these rules by a simple differential equation based on thermodynamics: $-Q' = RT^2(1/p' \partial(p' / \partial p) / \partial T - p' \partial p' / \partial T)$, where Q' is "partial heat of mixing" of 1 mole of a substance, p' its partial pressure over the mixt. and p' its satn. pressure.

at the temp. T . With the aid of this equation it is possible to calc. any change of min. and max. of the total pressure of a mixt. at any temp. if gas equations can be applied, at least approx., to vapors of liquids. A mixt. of FeOH and H_2O that does not obey Merriman's rule conforms strictly to this new equation. Conditions are set which are indispensable for the validity of Merriman's rule.

An extension of Ramsay and Young's boiling-point rule. T. S. WHEELER. *Phil. Mag.* 11, 441 (1931). —For single or pairs of liquids to which the integrated form of the van t Hoff isochore is applicable, the following general rule is deduced: "the reciprocals of the abs. temp. for which 2 given powers of the vapor pressure of a liquid or pair of liquids are in a const. ratio, satisfy a linear equation." Cox's rule (cf. *C. A.* 17, 2359) is implied in the Ramsay-Young rule which is a special case of the above. The rule was tested and found satisfactory for H_2O , H_2 , S , H_2 , and CH_3COOH for ratios from 2:1 to almost 1000:1. It was also found satisfactory for pairs of liquids and held for aq. NaOH solns. with more than 50% H_2O and for unsatd. NaCl and Na_2SO_4 solns. By plotting the reciprocal abs. temp. for H_2 and H_2O satisfying the equation $p_{\text{H}_2} = \sqrt{p_{\text{H}_2\text{O}}} \times 10^3$ it is possible to express the whole range of H_2 from -39° to 1300° . From the slope of the reciprocal temp. curve it is possible to calc. the ratio of the latent heats of 2 liquids.

Viscous flow and surface films. RONALD BULKLEY. *Bur. Standards J. Research* 6, 89-112 (1931). —Various fatty and mineral oils do not clog capillary tubes as small as 0.18-0.13 mm. in radius provided the liquids are properly filtered. The same statement applies to several other liquids if Pt and glass capillaries of radii 0.335 μ and 5.50 μ are used. The liquids retain their bulk viscosities up to at least 0.03 μ distance from a solid surface.

A new method of comparing viscosities of liquids by oscillating columns. G. SUBRAMANIAM. *Proc. 15th Indian Sci. Congr.* 1928, 86. —It is shown that in the case of a liquid contained in a U tube and making small and steady oscillations the rate of decay of such oscillations is related to the coeff. of viscosity by the expression: $\nu = \delta^2 \lambda / 2\tau$, where ν is the kinematic coeff., λ the logarithmic decrement, τ the free period and δ a const. which is assumed to be the same for the same tube and different wetting liquids. As the period of oscillations τ depends only on the total length of the liquid column and the inclination of the arms of the bent tube to the horizontal the above expression leads to a new method of comparing viscosities by noticing the log decrements of equal lengths of different liquids oscillating in a given tube. This relation has been verified in the case of 5 different tubes of 0.3-1.0 cm. diams. and 7 liquids, namely, water, alc., ether, xylene, benzene, turpentine and CS_2 , and the results (not given) are in perfect accordance with the theory.

The influence of the proximity of a solid wall on the consistency of viscous and plastic materials. III. R. K. SCHOFIELD AND G. W. SCOTT BLAIR. *J. Phys. Chem.* 35, 1212-8 (1931). —"A solid wall may modify the consistency of clay pastes at an appreciable distance from it but the available data are not sufficiently accurate to est. the thickness of the modified layer." A value of 8×10^{-3} cm. could escape detection because of exptl. uncertainties. Even if the effective thickness of the modified layer should amount to 20% of the radius of the tube, the construction proposed in the first paper (*C. A.* 24, 2022) closely approximates the mobility of the material in bulk.

Methods for determining the density of solid substances, particularly of inorganic salts. PRIMA WULF AND ALOIS HAIGL. *Z. physik. Chem., Abt. A*, 153, 187-209 (1931). —A series of d detns. on CsCl by the pycnometer method showed that, because of the impossibility of completely freeing the salt from occluded air, results are about 1% low. By using the flotation method, measurements of the d s of a series of salts were made with an av. error of only 0.2%. By the displacement method, the d s of salts too heavy to be detd. by flotation methods could be measured rapidly with an accuracy of about 1%.

Compressibility and pressure coefficient of resistance, including single-crystal magnesium. P. W. BRIDGMAN. *Proc. Am. Acad. Arts Sci.* 66, 255-71 (1931). —Measurements of compressibility at 30° and 75° are reported on 10 materials. All except metallic Mg crystallize in the cubic system so that measurements were made on cryst. aggregates. Mg crystallizes in the hexagonal system and single crystals were prepd. and the compressibility was measured parallel and perpendicular to the axis. The materials measured were NaP , BaF_2 , SrF_2 , CdF_2 , AlSb , CdTe , HgTe , TiN , TiC and Mg. The effect of pressure on the elec. resistance was measured for TiN , TiC and Mg.

Recrystallization and coloring of rock salt. KARL PRIZBRAM. *Sitzb. Akad. Wiss.*

Wien, Abt. IIa, 139, 255-70(1930), cf. C. A. 24, 1295.—The recrystn. time pressure-temp diagram was carried further. The recrystn. velocity, which was measured directly, often remains const. for several hrs. and then changes suddenly. Measurements were carried over a range from room temps., at which the van't Hoff coeff is valid, to 200°. The influence of weak radiation, formerly considered significant, is of no importance but, on the other hand, the arresting influence of strong radiation upon salt under high pressures was always confirmed. In regard to recrystn., melted crystals behave as if they contain more nuclei, which, however, show less tendency to grow, than the unheated salt. DON BROUSE

Recrystallization of pressed rock salt. GUSTAV ORTNER *Sitzb. Akad. Wiss. Wien, Abt. IIa*, 139, 271-4(1930).—Röntgen photographs show a marked crystal growth in rock salt after deformation by pressures of 5000 kg per sq cm. Molten crystals, on the other hand, display no grain growth after deformation. There is weak recrystn. in powder rock salt under high pressures. DON BROUSE

Dispersion in mixed crystals miscible in all proportions. G. TAMMANN and A. RUPPELT *Z. anorg. allgem. Chem.* 197, 65-89(1931).—Certain mixed crystals such as those of KCl-NaCl show turbidity, on cooling or on heating, at definite temps. A curve of beginning of turbidity can be obtained in either case which has a max. at some mol. fraction. Measurements of such mixed crystals were made for a considerable no. of salt pairs, the data from which together with others obtained from the literature are collected in tabular form and by graphs. The linear rate of turbidity increase was measured by cooling the salt pair in a glass tube after which on cutting the tube at some spot, turbidity increases into the clear portion of the mixed crystal. A very short range in mol. fraction for turbidity exists for LiBr-NaBr. The presence of water greatly affects the amt. of dispersion, the influence of which quickly disappears at about 100°. The salt pair KPO_3 - $NaPO_3$ shows a eutectic in the diagram for heating not obvious in the cooling curve. Little turbidity is shown. Mention is made of Na and K-contg. feldspars. Opalescence can be obtained by controlling either rate of cooling or degree of moisture present. A table is given relating the dependence of miscibility on the relative differences in lattice parameters. If this difference is less than 0.115 A. U., unlimited miscibility exists while miscibility practically disappears at, or above, 0.144 A. U. R. H. LAMBERT

Diffusion of sodium in rock salt. E. REXER *Physik. Z.* 32, 215-6(1931).—Na was placed in small holes in rock salt crystals, and diffused into the crystals at higher temps. The distribution and size of the particles after diffusion were studied with the ultramicroscope. C. E. P. JEFFREYS

Melting point curves of the monobasic fatty acids. ANNIE MILLICENT KING and WM. EDW. GARNER *J. Chem. Soc.* 1931, 578-80.—The entropy change on crystn. can be accurately represented by the equation $Q/T = 0.002693n - 0.0061 + 0.00475 \sum_{i=1}^n \log_{10} n/(n-2)$, where n is the no. of C atoms. In order to make a calcn. of the m. p. of an unknown acid, it is necessary to be able to det. a value for Q . The manner in which the heat of crystn. varies with the no. of atoms in the C chain has not been interpreted theoretically, so that it is only possible to employ an empirical relationship. For acids having more than 12 C atoms, Q increases at the rate of 2.060 cal./g. mol. for every $2CH_2$ groups added to the chain. C. J. WEST

Lattice energy and the state of combination. WILHELM KLEMM *Z. physik. Chem., Abt. B*, 12, 1-32(1931), cf. C. A. 25, 11.—The quotient of the exptl. lattice energy by the normal lattice energy, the latter being calcd. on the assumption that the lattice is formed of ions of the rare gas configuration, gives a more useful comparison than their difference when anions and cations are very different in size or for unsatd. compds. A similar quotient for gas mol. energy is described. The stability of compds can be predicted from these terms. The transformation from coordination to mol. compds takes place in a homologous series as the cation increases in wt. or as the valence of the cation increases for solids and from ions to atoms in gases. No distinct boundary appears to exist between these types of union from an energy relationship. Most calcns. were made on halogen compds. R. H. LAMBERT

The arrangement of the microcrystals in white tin deposited by electrolysis. HIDEKI HIRATA, HISAJI KOMATSUBARA and YOSHIO TANAKA *Anniversary Vol. Dedicated to Masumio Chikashige* (Kyoto Imp. Univ.) 1930, 261-73.—The cryst. structures in electro-deposited specimens of white Sn, obtained under various conditions, were examd. with x rays, by the so-called "transmission method." The cryst. structure is influenced by at least 2 factors: c. d. and the concn. of the electrolyte. Smaller c. d. and lower concn. of the electrolyte are the more favorable conditions for the growth of large

crystals. When these 2 factors are comparatively large, the microcrystals of white Sn tend to be deposited in a fibrous form. The microcrystals have a tendency to be electrolytically deposited with the normal to their octahedral faces arranged parallel to a definite common axis. ALLEN S. SMITH

The crystal structure of the inert gases. II. Krypton. A. NASINI and G. NATTA. *Atti accad. Lincei* 12, 141-7 (1940). The method of powders was used as with Ne (C. A. 24, 3410-4168). The crystal structure of Kr was detd. at the temp. of liquid N₂. It is a face-centered lattice having 4 atoms with side 5.78 Å U., vol. 193×10^{-24} cc., $d = 2.83$. The radius of the atom as detd. by x rays was compared with that detd. by viscosity measurements for Ne, Ar, Kr, N₂, CO, NH₃, H₂S, HBr, Xe, PH₃, HI, the ratio of the value found by these 2 methods varying from 1.28 to 1.35, indicating that these 2 methods of measurement are concordant. A. W. CONTIERI

Structural analogies of binary alloys of transition elements and zinc, cadmium and aluminum. WALTER EKMAN. *Z. physik. Chem., Abt. B*, 12, 57-78 (1931).—Powder x ray diagrams were obtained of phases of the systems Fe-Zn, Co-Zn, Ni-Zn, Rh-Zn, Pd-Zn, Pt-Zn and Ni-Cd. All correspond to the interference photograph of γ -brass. These phases represent a valence electron concn. of 21-13, on the assumption that transition elements may be regarded as of zero valence. The homogeneous region for an Fe-Zn phase of this type lies between 19 and 23 atom % Fe. The Co-Zn phase and the Ni-Zn phase are homogeneous between 15 and 22 atom % Co and 15-19 atom % Ni, resp. Co-Zn shows a phase similar to β Mn in structure. A phase analogous to β -brass was found in the Co-Al system. Calcn. of at. vol. indicates that contraction accompanies phase formation. FRANK URRAN

Hydroxide systems in iron oxide colors. HANS WAGNER. *Kolloid-Z.* 54, 310-4 (1931), cf. C. A. 25, 221.—X ray photographs showed the following structures for various conch. colors: (1) bright ochre 1 Fe₂O₃, goethite lattice; (2) Mars yellow contg. CaCO₃ and CaSO₄, CaCO₃ lattice; (3) Mars yellow free of chalk, no lattice; (4) Tuscan sienna, goethite lattice; (5) Fe(OH)₃, from FeCO₃, weathering, goethite lattice; (6) Fe(OH)₃, from sericite weathering, mica lattice. Ochres or citron yellows are characteristic of colors composed of monohydrate. Microscopic examn. reveals 1 and 3 contg. very fine needles with a strong tendency to agglomerate, 4 consists of coarse grains, splinters and angular particles which are transparent in a bright field and shine in polarized light. 2 and 3 resemble 4 more closely than 1. Ochres may be classified according to the adsorption of basic dyes, such as Brilliant Green, by the other substrate, usually a mixt. of Al₂O₃, SiO₂ and H₂O. German ochers adsorb the dye irreversibly (ebemosorption) and the substrates are Al₂O₃-SiO₂ gels, French ochers adsorb the dye reversibly (lysoorption) and the substrates are kaolins. The classification is not absolute and transition types of ochre exist. ADRIAN FLEISCHER

The structure of hydrogen sulfide and hydrogen selenide. G. NATTA. *Nature* 127, 129 (1931).—A cubic lattice was found with face-centered distribution for S and Se atoms. The d of H₂S was 1.166, H-Se 3.45. The side of the cell (a) was found to be 5.778 Å U. for H₂S and 6.020 for H₂Se. Considering both as ionic compds., the most probable space group is Oh^h , a fluorite type shown by Li₂S. FRANK MARESH

Spinel. III. The titanates of cobalt and of zinc. L. PASSERINI. *Gazz. chim. Ital.* 60, 937-62 (1930). cf. Natta and P., C. A. 23, 4649, 24, 1311, P., C. A. 23, 4167, 24, 4439.—The present paper is a continuation of studies of compds. that crystallize in the cubic system with a lattice structure of the spinel type. Co orthotitanate, Co₂TiO₄, not heretofore described, was prepd. by heating rapidly to 900° an intimate mixt. of TiO₂ and Co(NO₃)₂ (2 mols). The product was a dark-green, pulverulent substance. A little CoTiO₃ (cf. *Bull. soc. chim.* 15, 194 (1893)) may be formed. Zn₂TiO₄ was prepd. by the method of Levy (*Compt. rend.* 107, 421 (1888)). Traces of ZnTiO₃ are also formed. Its phys. properties do not agree with those described by Lévy; e. g., it is a white, pulverulent substance, with much higher d than the 4.16 of Lévy. Both Co₂TiO₄ and Zn₂TiO₄ were examd. by the powder method, with K α and K β radiation from an Fe anticathode (Philips tube). Both titanates crystallize in the cubic system, with a lattice structure of spinel type (cf. 16c, 32b, spatial group O_h 7-). For Co₂TiO₄, $a = 8.420 \pm 0.005$ Å U., $V = 596.94 \times 10^{-24}$ cc., calcd. $d = 5.12$ (for a cell contg. 8 mols). The Co₂TiO₄ was not sufficiently pure to det. the exptl. d . For Zn₂TiO₄, $a = 8.410 \pm 0.005$ Å U., $V = 594.82 \times 10^{-24}$ cc., calcd. $d = 5.43$. C. C. D.

Comparison of the crystal structures of an addition and a substitution compound. EDUARD HERTEL and KURT SCHNEIDER. *Z. physik. Chem., Abt. B*, 12, 109-14 (1931), cf. C. A. 24, 2924.—The crystal structures of aniline picrate (addn. compd.) and 2,4,6-trinitrodiphenylamine (substitution compd.) showed an extensive analogy, according to x ray analyses. They agreed with respect to the spatial group and 2 main identity

periods. The difference in the third period may be due to a sp orientation of the mol in this direction, and to spatial requirements of the H_2O mol, by which addn and substitution compds differ. FRANK URBAN

Crystals of 3,4,3',4',6'-pentamethoxydiphenylmethane-2-carboxylic acid. E. LENGYEL. *Acta Sci Univ Francisco Josephinae, Acta Chem, Mineral., Physica* 1, 68-71(1929). Crystallographic description of rhombic crystals. S S DE FINALLY

X-ray examination of the crystal structure of resorcinol. A N SARKAR. *Proc 15th Indian Sci Cong* 1928, 92. — X ray examns of crystals of catechol, resorcinol (I) and hydroquinone were undertaken with a view to confirm the structure of the benzene nucleus detd by Bragg from the examn of $C_{12}H_6$. The data obtained from the powder photographs of crystals of I were insufficient for the detn of mol structure, but it was found that the space group is C_{2v}^{12} and the mols are asym. From certain observed accidental interference minima, not characteristic of group C_{2v}^{12} , certain relative arrangements of the 4 types of mols in the unit cell were deduced. S's deductions as regards space group and no per cell confirm Bragg's conclusions, but S suggests a different arrangement of mols. E J C

Polymerization in crystal lattice—crystal structure of trinitroresorcinol and trinitrophenol. EDUARD HERTEL AND KURT SCHNEIDER. *Z physik Chem., Abt B*, 12, 139-50(1931), cf. *C A* 24, 2921. — Picric acid has 4 double mols in its unit cell (cf Bridg and Möller, *C A* 23, 5368). The entrance of a second Off group as in styphnic acid (I) causes the formation of 2 triple mols in the unit cell but the entrance of a third Off as in 2,4,6 trinitrophenol (II) does not continue the "polymerization". I and II are trigonal, probably space group C_{3v}^1 , but possibly of D_{3d}^2 . Their crystal structures are almost identical. The unit cell of each contains 6 formula weights and has the dimensions for I: $a = 12.7$, $b = 22.0$ and $c = 19.0$ Å. and for II: $a = 13.4$, $b = 23.6$ and $c = 9.6$ Å. The structure of each is built up of 2 units of $3 C_{12}H_3N_3O_7$ or $CaH_3O_7N_3$ groups in the unit cell. Test for piezoelec effects with both was neg. In spite of the far-reaching similarity of the cryst structure of both compds they form no series of mixed crystals, the phase rule diagram shows only a eutectic. O F S

The production of residual double refraction by pressure in certain glasses at atmospheric temperature. F. C. HARRIS. *Phil Mag* 11, 747-8(1931). — A uniform stress was applied for a period of 3 years on glass disks contg 10, 47, 64 and 79% PbO. The first 2 glasses had a pos and the last glass a neg stress optical coeff. A definite residual double refraction was observed for the last 3 glasses. This residual refraction disappeared 1 month after the stress was removed. The time necessary to produce the effect decreases with increase in the PbO content, the effect increases with increase in the PbO content. In no case was an increase in the residual double refraction noted during the time the load was continued. ARTHUR LEISCHER

The birefringence of safrole. L'AUTHENIER AND BART. *Compt rend* 192, 352-3(1931). — No appreciable residual birefringence was found for com or carefully purified safrole, either in a const or alternating field. The Kerr const for pure safrole is $K = 0.82 \times 10^{-7}$ for the green fig line $\lambda = 580.0$ Å. P. W. LAIRD

The nature of "triatomic hydrogen." EGON HIEDEMANN. *Z. physik. Chem., Abt. A*, 153, 210-40(1931). — By passing an electrolytic ring discharge through H_2 in a glass reaction vessel a gaseous product is formed that is easily condensable in liquid air, and that shows the chem characteristics ordinarily ascribed to the so-called "triatomic hydrogen". Each vol of it, on reacting with Na gives 2 or 3 vols of H_2 . It reduces WO_3 to WO_2 , reduces $KMnO_4$ soln, decolorizes indigo blue, gives a yellow coloration with Nessler's reagent that is characteristic of NH_3 , and causes a brown coloration on lead acetate paper. These same chem properties are characteristic of the mixt of Si hydrides obtained from the reaction of Mg silicide and HCl . After a crit review of all previous papers, it is concluded that all of the "triatomic hydrogen" reported in the literature has been either H_2S or a mixt of Si hydrides, in some expts. both gases probably have been present. P. H. EMMETT

The electrical resistivity of silicon. CH BEDEL. *Compt. rend.* 192, 802-4(1931). — Fast detn of the elec. resistivity of Si have been vitiated by the difficulty of obtaining compact specimens of pure Si, and of establishing good elec contact with the specimen. These difficulties were overcome, and the rv resistivity of pure Si contg only 0.1% Fe was found to be 0.267 ohms per cm. The rv resistivity of impure Si contg 1.4% Fe was 0.003, and of a ferro-silicon contg 8.5% Fe, 0.0035. For poor contacts involving large surfaces, the resistance changes when the direction of the current is reversed, it decreases if the current intensity is increased, but increases with the duration of flow of the current; and it diminishes with increased pressure. Poor contact has been at-

reduced to a thin layer of NaCl or of adsorbed gas. However, it persists after treatment with HF and the deposition of an adherent coating of Cu or Ag. On the contrary, the electrolytic deposition of Zn or Ni by a current of 0.2 amp. provides good contact. Good contact was obtained also by depositing Cu on the Si by immersion in a soln. of Cu^{+2} , then coating the Cu with a Sn amalgam, Hg:Sn = 15:2, and finally applying the vacuum dem. The samples were held under pressure between Cu electrodes during the resistance dem. R. H. LOUGHRAN

Sublim. free from cerium. M. PIERCE. *Compt. rend.* 192, 684-6, 1931).—“Cerium chloride (CeCl_3) can be prepd. at 1350° in a boat of graphite by the action of H_2S on ceria. This compound is remarkably resistant to the action of heat; it melts only above 2100° and it volatilizes only at 2200° under a vacuum of about $1/1000$ mm. Hg. It resists equally the action of metals and the non-oxygen salts. CO_2 causes a special oxidation at 700° in which Ce is transformed into CeCl_3 , S eliminated, and CO_2 reduced to CO .” DON BAILEY

Insulation of Pyrex glass after heating in vacuo. J. H. MITCHELL. *Phil. Mag.* 11, 74-85, 1931).—Pyrex glass on heating in vacuo shows a rapid increase in cond. at temps. from 55° to 80° , the exact temp. varying with the specimen. At 250° to 300° a decrease and at 350° to 400° a marked increase in the cond. were found. On cooling, the cond. again decreases but is much higher than the original value before heating. Cooling to liquid-air temp. did not decrease the cond. Washing the surface of the glass with HF restored the initial low cond. Heating in air at 500° also decreased the cond. but not to the initial value. The glass treated in this way did not increase in cond. on reheating in vacuo. The permanent increase in cond. is a surface effect, probably due to the chem. changes in the outer layers of the glass. ARTHUR FLAJOULX

The adsorption of ions on a surface film. R. S. BRANTY. *Phil. Mag.* 11, 44-53, 1931).—Math. The potential due to a uniform plane array of dipoles oriented perpendicularly to the surface is calcd. The results are extended for fatty acids in 0.01 N HCl. ARTHUR FLAJOULX

Adsorption of gases by glass walls. VIII. Hydrochloric acid gas. M. CASSIDY AND E. MURPHY. *Trans. Am. Chem. Soc.* 53, 14-37, 1931). *J. C. S. A.* 24, 2837.—Adsorption of HCl by the glass walls of the flask used to det. the mass of the normal l. and of the app. used to det. compressibility gave -1.74×10^{-4} g. at 700 mm. for a 1-l. flask, as the correction to be applied. The tests confirmed the McRae theory of both absorption, or solid soln. of gas on the glass surface, and adsorption on the surface. The course of both processes can well be expressed by the Freundlich adsorption equation. Results of others are criticized. E. M. STROVER

The adsorption of certain vapors by activated charcoal. III. J. N. FRANCES AND A. L. TAYLOR. *J. Phys. Chem.* 35, 1091-1103, 1931). *J. C. S. A.* 24, 2890.—The adsorption of ethyl, n-propyl, n-butyl and tertiary butyl chlorides on charcoal was detd. from 0° to d.-compn. temps. The stability decreases with increase in mol. complexity. The amt. of adsorption in general increases with mol. wt., with branched-chain compounds less than with straight chains. The heats of adsorption are less when calcd. from the curves than the exp't. values. The distribution of potential and d. are calcd. according to Polanyi. CARLOS M. EVANS

Adsorption of sugars by animal charcoal and vegetable decolorizing carbon. N. TAKENAKA. *J. Sci. Chem. Ind. Japan* 30, 70-83, 1927). *Memo. Facult. Sci. Exp. Waseda Univ.* 1930, No. 7, 125-6.—The amt. of sucrose adsorbed from an aq. soln. by animal charcoal increases to a max. at a concn. of 0.2 mols. per l. Freundlich's equation applies to the dil. soln. The amt. of adsorption is directly proportional to the charcoal added. Grape sugar is also adsorbed by animal charcoal. Sucrose is adsorbed in appreciable amt. by vegetable decolorizing carbons. L. W. T. CRONIN

Effect of gelatin and salts on Congo red. HERBERT L. DAVIS AND JOHN W. ACKERMAN. *J. Phys. Chem.* 35, 972-87, 1931).—Gelatin exhibits an apparent protective action against pptn. of salts on solns. of Congo red, due to a decreased stability and decreased dispersion of the dye. FRANK URBAN

The distinguishing features between adsorption by a substance in a hydrated and a dehydrated condition. M. P. VENKATARAMA IYER. *Proc. Ind. Acad. Sci. Chem.* 1928, 132-3.—It has been found from measurements that benzoic, cinnamic, salicylic and stearic acids, which are generally dehydrated, show hardly any change in the f_{H} values or the elec. charge when kept in contact with neutral salt solns., whereas stearic acid which is highly hydrated in the colloidal state shows marked changes in the f_{H} values and elec. charge when similarly treated. Dehydrated substances like the above adsorb their constituent ions very strongly, whereas hydrated substances adsorb any ions without showing any strong preference for the constituent ions. Irreversible coagulation and

finely divided org substance immiscible with water is introduced into a coazervate system the org droplets are taken inside the coazervate drops. Only those droplets so taken up manifest a shift in the field of the elec current. The direction of this shift is independent of the nature of the org fluid, of the pH or the presence of neutral salts, but is detd by the cataphoretic charge of the coazervate drops. If the latter are pos. the shift is toward the neg electrode, and if neg the shift is to the pos electrode.

S. MORGULIS

The dialysis of sodium bicarbonate solutions. F. CANALS AND J. DAUBAIN DELISLE. *Bull soc chim bel* 12, 1165-61(1930)—The CO_2/Na value of a bicarbonate soln. varied from 1.98 to 1.74 during 2 hrs' dialysis.

C. G. KING

Ultrafiltration as a test for colloidal constituents in aqueous and non-aqueous systems. J. W. MCBAIN AND S. S. KISTLER. *J. Phys Chem* 35, 130-6(1931)—Cellophane membranes show unusual properties in dialysis, they retain all but the smallest colloidal particles in diffusion. It has been possible to demonstrate with these membranes, that colloidal constituents are commonly present in non aq. solns. of the electrolytes $AgNO_3$, NH_4I , $AgBrO_3$, CdI_2 , etc., which exhibit anomalies in elec. cond. and osmotic behavior. The non aq. media for $AgBrO_3$ and $AgNO_3$ are $LiNH_2$ and piperidine, resp.

H. R. MOORE

Some chemical reactions of colloidal clay. RICHARD BRADFIELD, *J. Phys Chem* 35, 358-73(1931)—The base-exchange phenomena and the cation-exchange reactions of colloidal clays were studied experimentally. Methods are developed for measuring the solid acidity of li clays, and the significance of titratable acidity of electroanalyzed clays is pointed out. Changes in the phys. properties of clays may be correlated with changes in electrophoretic potential.

H. R. MOORE

Thermo-sensescence effect exhibited by gold sols at elevated temperatures and aging at room temperature. EMLYN JONES AND W. C. M. LEWIS. *J. Phys Chem* 35, 1168-73(1931)—In general no sol is in a suitable state for exam. until it has been dialyzed for about 4-5 days and has then been subjected to about 3 hrs' heat treatment. Thermo-sensescence may not be attributed to the further removal of impurities, but is due to the contraction of the sol particles as a result of raising the temp. The aging effect, however, may well be due to ionic adsorptions and desorptions.

FRANK URBAN

Studies on the formation of silver sol prepared by dispersion in the electric arc. P. S. MACMANON AND S. C. VARMA. *Proc 15th Indian Sci Congr.* 1928, 136—Attempts were made to det. the properties of finely dispersed Ag. The deposit obtained by Furth (C. A. 18, 2449) from an elec. arc struck in air between Ag electrodes gave a sol on treatment with water. However, this deposit is by no means pure Ag, it contains large quantities of oxide and nitrate. When the dust is ignited, brown fumes are given out in perceptible quantity. The dust was prepd by striking the arc between pure Ag electrodes in either air, O or N. Nitrate was found to be produced in O contg. as little as 0.2% N or in N contg. a little O. In pure N practically no deposit was formed, and it is probable that if the gas were chemically dry there would be no formation of dispersed Ag whatever. The conditions under which the sol was produced by the addn. of water to the powder were studied at length, but the results are not given. The properties of the solid powder were also investigated. There is evidence that this powder exposed to air or O takes up a further small quantity of O, indicating the oxidation of finely dispersed Ag.

E. J. C.

Colloidal platinum and its behavior as a typical acidoid sol. S. W. PENNYCUICK. *J. Am. Chem. Soc.* 52, 4621-35(1930)—Evidence is presented for the behavior of colloidal Pt as a typical neg hydrophobe or acidoid sol. Deductions are made on the mechanism of retention of acid on the surface, and the fraction of surface covered by acid preferentially. $H_2Pt(OH)_2$ appears to be combined at the surface with oxides of Pt, held together by the usual forces of chem. affinity. Ionic replacement of H^+ by Ba^{++} , for example, can occur, and the surface oxide of the sol $[xPt_2yPtO_3zPt(OH)]^{2-}$. $2H^+$ which is acidic may combine to form simple and complex salts. An extensive study is made of coagulation phenomena which occur freely with bases and salts but are restricted with acids. It is shown that HCN increases the charge carried by the colloidal particle, and likewise displaces the surface $H_2Pt(OH)_2$.

H. R. MOORE

Preparation of lead and lead phosphate sols. VICTOR CORMAN. *Rev gen colloides* 8, 337-57(1931)—Various methods of prep. Pb and Pb phosphate sols are given. The effect of several stabilizing agents (gelatin, glucose, $Na_2S_2O_3$, and various salts) was investigated. The particle size was detd from the rate of deposition and by means of the ultracentrifuge. Striking lowering of the pH occurs in the prepn. of colloidal Pb phosphate.

FRANK URBAN

Tungstic acid hydrosol. A. R. NORMAND AND M. C. NUTHANNA. *Proc. 15th*

Indian Sci Cong 1928, 137—It has been stated that tungstic acid sol is indifferent to acids, salts and alies. But a fairly sensitive sol has been obtained by peptizing the tungstic acid both with an acid and an alkali. The sol of tungstic acid prepared in either of these ways is negatively charged and has the characteristic properties of an emulsoid sol. The order of coagulating power of the chlorides of alkali metals (determ. with the aid of the photoelec. cell) is $\text{PbCl}_2 > \text{KCl} > \text{NaCl} > \text{LiCl}$. The protective influence of the neg ions is $\text{Cl}^- < \text{Br}^- < \text{I}^-$. E J C

The stabilization of blue cupric hydroxide. HARVEY A. NEVILL AND CHARLES T. OSWALD. *J Phys Chem* 35, 60, 72 (1931). The blue substance obtained by the action of alkali on a cupric salt is shown by x ray analysis with the diffraction pattern obtained by λ_{Cu} to be $\text{Cu}(\text{OH})_2$ or the hydrated oxide $\text{CuO} \cdot \text{H}_2\text{O}$. A 5% gelatin soln stabilizes this colloid. H R MOORE

Ion interchanges in aluminum oxychloride hydrosols. ARTHUR W. THOMAS AND THOMAS H. WHITEHEAD. *J Phys Chem* 35, 27, 47 (1931).—A technique for the prepn of Al oxychloride hydrosols of colloidal dimensions is presented. The following set of properties of the sols has been established: (1) Neutral salts of the type K_2SO_4 , KOAe , KCl and KNO_3 increase the pH of the sols from 5 to 6.8 almost to 7.0 approx. (2) Neutral salt titration of AlCl_3 sols of different haucities leads to a rate of change of pH which is a function only of the no. of hydroxo groups, viz. the series $\text{Al}(\text{OH})_2\text{Cl} \rightarrow \text{Al}(\text{OH})\text{Cl}_2 \rightarrow \text{AlCl}_3$. (3) Several anions vary in their "penetration power" toward the basic Al ion, leading to an order for anions. The authors' extension of Werner's theory adequately explains the formation of the sols, their decrease of acidity on aging, increase on heating and the effect of heat on the soly of Al hydrates. H P MOORE

Surface processes on coagulating precipitates. II. The mechanism of adsorption in electrolytic solutions. LUDWIG IMAN. *Z physik Chem, Abt A*, 153, 127-42 (1931), cf C A 24, 2973. The process of the adsorption of ions on the surface of aging suspension of Ag halides was examd. The changes with lapse of time of the adsorption values are different and depend on the soly of the compd. of the corresponding ion with the oppositely charged constituent of the lattice. This different behavior leads to conclusions about the process of adsorption. At certain areas of activity on the surface of the aging ppt, the lattice is loosened because of the tendency to crystallize. A process of interchange is possible with the adsorbed ions which are in the immediate vicinity of the lattice. A HIRSCH

Equilibrium phenomena in coagulation of colloids. E. F. BURTON AND MAY AVERY. *J Phys Chem* 35, 44, 59 (1931).—An ingenious optical study is made of the scattered and transmitted light from aq soln of mastic, Au, and As_2S_3 . Variations in the d. of scattered light are recorded, as well as a curious zonal effect in the coagulation of As_2S_3 . Coagulation by stages is observed on adding traces of electrolyte to the sols, and evidence is obtained for equal states in the coagulation process. The As_2S_3 sol probably has a constitution of the type $(\text{nAs}_2\text{S}_3 \cdot \text{HS})^- + \text{mH}^+$, the H ions diffusing as the outer layer of the Helmholtz double layer. H R MOORE

The mechanism of the coagulation of sols by electrolytes. I. Ferric oxide sol. HARRY B. WEISER. *J Phys Chem* 35, 1, 26 (1931).—The coagulation process is studied by following the change in compn. of Fe_2O_3 sols by adding pptg. electrolytes stepwise to the sol. Cl^- is released in the reaction $2[\text{xFe}(\text{OH})_2 \cdot \text{yFeOCl} \cdot \text{FeO}]^+ + 2\text{Cl}^- + \text{K}_2\text{SO}_4 \rightarrow [\text{xFe}(\text{OH})_2 \cdot \text{yFeOCl} \cdot \text{FeO}]_2\text{SO}_4 + 2\text{KCl}$, and the increase in $[\text{Cl}^-]$ may be followed potentiometrically. In micelles of this type thousands of equivs may exist for each free charge. Several titration curves are obtained with sols of 3 distinct types and with inorg. salts of different valence types as exemplified by K_2CrO_4 , $\text{K}_2\text{C}_2\text{O}_4$, and $\text{K}_4\text{Fe}(\text{CN})_6$. A mechanism of coagulation is proposed from these data, postulating positively charged micelles as a consequence of the hydrolysis. These micelles vary in size, compn. and charge depending on conditions of hydrolysis. H R MOORE

The influence of electrolytes on the coagulation of ceric hydroxide hydrosol heated to different temperatures. A. R. NORMAND AND M. C. MUTHAYYA. *Proc. 15th Indian Sci Cong 1928, 135*.—The coagulating powers of the chlorides of alkali metals and metals of the alk. earths on pos. $\text{Ce}(\text{OH})_3$ hydrosol, heated to different temps., were studied. The time required for each electrolyte to coagulate the sol was followed by means of a K photoelec. cell. The order of the coagulating power of the electrolyte was found to change with different degrees of dehydration of the sol. The order of the coagulating power of the electrolytes is: (1) With sol dialyzed at 28° , $\text{LiCl} > \text{NaCl} > \text{KCl} > \text{RbCl}$; (2) With sol dialyzed at 28° and heated to 70° , $\text{LiCl} > \text{NaCl} > \text{RbCl} > \text{KCl}$; (3) With sol dialyzed at 28° and heated to 100° , $\text{KCl} > \text{RbCl} > \text{LiCl} > \text{NaCl}$. Thus the coagulating power of the different electrolytes depends upon the temp. condition of the sol. F J C

The vapor-adsorption capacity of silica gels as affected by extent of drying before wet-heat treatment and by temperature of acid treatment and activation. HARRY N. HUBBINS AND A. L. ELDER. *J Phys Chem* 35, 82-82 (1931).—Eleven SiO_2 gels showed under 1 mm benzene adsorption on aging for a few years. Four SiO_2 gels were heated at 100° without causing serious decreases in their benzene adsorption capacities. Increase in temp. of acid treatment of a gel from 30° to 100° increases the porosity of the gel. 10% adsorption of benzene by weight was obtained with a gel. Chalky gels have high total adsorption capacities at high partial pressures but are not efficient for adsorption at low partial pressures. A. L. ELDER

Adsorption of acids by silica. M. P. LAXHANI. *Proc. 15th Indian Sci Cong* 1928, 134. In view of the controversy between Joseph and Mukherjee (*C. A.* 21, 544) it appeared that the method of prep. SiO_2 detd. whether it would adsorb acid or not. CaSO_4 and H_2SO_4 mixts. were distd. at low temp. The gas was collected under water. The ppt. of SiO_2 thus formed and dialyzed for 7 days did not adsorb HCl or H_2SO_4 , as indicated by analytical elec. cond. and pH detns. E. J. C.

Adsorption by silica gel from binary mixtures of liquids. BASRUR SANJIVA RAO AND H. N. CHANNARASAPPA. *Proc. 15th Indian Sci Cong* 1928, 133.—Silica gel particles have a water envelope, and B. S. Rao (Thesis, London University, 1926) has shown that selective adsorption from binary mixts. of liquids depends on the energy of the interface that this water envelope forms with each of the 2 liquids. The interface between water and CCl_4 has the same energy as that between water and CS_2 . R. and C. did at 30° adsorption by silica gel from mixtures of CCl_4 and CS_2 in different concns. and in conformity with the above view, found no selective adsorption. Detns. were carried out in 2 ways: by a static method in which the gel was treated with liquid mixts. in a thermostat and by a dynamic method in which air bubbled through the liquid and passed over the gel in a closed system, the circulation of air being effected by a double-acting pump of special design, working in an air thermostat and consisting of 2 pulsating Hg columns and 4 Hg valves. A Pulfrich refractometer was employed for the analysis of the mixts. E. J. C.

Alcohol of silica. BASRUR S. RAO AND K. G. DOSA. *Proc. 15th Indian Sci Cong* 1928, 133.—Adsorption by silica gel is ordinarily dependent on the water envelope in it. Attempts have been made to obtain gel with an alc. envelope. Expts. on the complete replacement of the water by alc. carried out by Graham, by Nannhauser and Patrick, and by Luth and Purse gave contradictory results. R. and D. extd. with hot alc. silica gel (placed in a Jena glass thimble) in a Soxhlet app., the alc. in the flask being kept anhyd. by addn. of metallic Ca and BaO. On prolonged treatment the water in the gel was found to be about 1% by org. combustion methods suitably modified. Replacement of water was also attempted by a dynamic method in an app. in which air was continuously circulated in a closed system through anhyd. alc. and the gel. Adsorption curves (not given) for silica gel and water-alc. mixts. indicate that replacement of water becomes progressively difficult with decrease in water content of the gel. Complete replacement of water, though it may not be impossible, is extremely difficult to effect. E. J. C.

The density of water adsorbed on silica gel. DWIGHT T. EWTING AND CHARLES H. SPURWAY. *J Am Chem Soc* 52, 4635-41 (1930).—Values obtained for the d. of water adsorbed on SiO_2 gel were secured by vol. measurements, a gas dilatometer with He as the inactive gas was used. For small quantities of water up to 4.36%, the measurements show that the d. of water adsorbed at 25.02° is greater than the d. of liquid water at the same temp. H. R. MOORE

The growth of lead crystals in silica gels. ROBERT TAFT AND JESSE STARECK. *J Chem Education* 7, 1820-36 (1930).—Pb crystals were grown in SiO_2 gels. The following exptl. factors were varied: nature of the reducing agent, concn. and compn. of the Pb salt, pH of the gel, and the viscosity of the intercellular liquid present in the gel pores. By using a device equiv. to a voltaic cell, Ag or Cu could be used as reducing agent. Numerous photographs illustrate the crystal growths. ALBERT L. HENVE

The physical state of water bound by organic colloids and by the tissues. N. MARINESCO. *Compt rend soc biol* 103, 872-5 (1930). *Physiol Abstracts* 15, 274.—The hydrophilic coeff. (no. of g. of water electrophoretically satd. and adsorbed per g. of colloid) is 15 with dog hemoglobin, 11-13 with egg albumin and 8-10 with gelatin. The coeff. is lowest at the isoelec. point. The micelles of a hydrophilic colloid are formed mainly of water, of which only a small part is free in an animal organism. The water in a colloid is not only fixed but compressed. During swelling of colloids heat is given out, which is a measure of the compression. G. G.

Properties of protein-cellulose membranes. LÉON VELLUX AND JEAN LOISELLEUR.

Compt rend 192, 306-8(1931) —The prepn of these membranes, which consist of solid sols. of proteins in cellulose esters, is given in *C A* 25, 1723, 1861. That the membranes are very homogeneous is shown by their even staining with dyes. When treated with a soln of an Au salt and a reducing agent the Au is pptd in colloidal form in the membrane which then appears red brown by reflected and deep blue by transmitted light. Each component of the membrane retains its particular properties. The cellulose ester being insol in H_2O and inert, plays a purely mechanical or static part, but the protein through its activity confers on the membrane properties that make it comparable to a certain degree with natural cell membranes. I. E. GILSON

Certain observations on the Donnan equilibrium. A. RONCATO *Boll. soc. ital. biol. sper.* 5, 1103-6(1930) —Isoclec gelatin was dissolved in 1% concns in 0.01 *M* to 0.001 *M* solns NH_4Cl corresponding to pH 5 to 6.5 and placed in collodion sacs. The H ion concn of the internal liquid was greater than that of the external, with 0.001 *M* NH_4Cl the difference in pH between the external and internal liquid was 0.40; with concd solns of NH_4Cl approaching the pH of isoclec gelatin (4.7), this difference was not so marked. The membrane potential found corresponded to the calcd. theoretical value. If $M/100,000$ to $M/1,500,000$ HCl was used instead of NH_4Cl , although the pH values of the solns were similar the difference in pH on both sides of the membrane was markedly less than that of the corresponding NH_4Cl soln. The supposition is made that NH_4OH and HCl , the products of hydrolysis of NH_4Cl reciprocally influence each other, and being simultaneously present change the Donnan equil. which would have been reached had only one of these products been present. The production of HCl in the stomach might be interpreted as the result of a concn. of HCl in the interior of the gastric cells brought about by a state of Donnan equil. reached in the presence of NH_4Cl . PETER MASUCCI

Adsorption of ions and sols by freshly prepared precipitates and its influence on the formation of Liesegang rings. II. A. C. CHATTERJI AND S. C. VARMA *Z. anorg. allgem. Chem.* 196, 247-56(1931) cf. *C A* 23, 5380 —No adsorption of CrO_4^{2-} by $PbCrO_4$ takes place during formation of the latter. However, adsorption of $PbCrO_4$ sol by well-washed $PbCrO_4$ occurs. This adsorption, in the presence of gelatin, is a function of the gelatin concn., it approaches zero at higher gelatin concns. This explains why it is difficult to prep. $PbCrO_4$ rings in such gelatin gels. FRANK URBAN

Evidence in favor of the existence of silver chromate in gelatin in the colloidal condition. Electric conductivity of silver chromate in gelatin. A. C. CHATTERJI AND S. C. VARMA. *Proc. 15th Indian Sci. Cong.* 1928, 142-3 —Bolam and MacKenzie (*C A* 20, 2772) contest the conclusion of Chatterji and Dhar (*C A* 19, 3191) that Ag_2CrO_4 exists in the colloidal condition when pptd. in gelatin and maintain that it exists in the ionic state. Expts. were undertaken to det. accurately the elec. cond. of Ag_2CrO_4 of different strengths in gelatin of varying concns at 35°. The Ag_2CrO_4 was produced by the addn. of equiv. quantities of $AgNO_3$ and K_2CrO_4 . The results (not given) confirm the earlier conclusions. From a study of the results of B. and M.'s e. m. f. detns. C. and V. point out that, except in a very few cases, not more than 40% of Ag exists in the ionic condition. Expts. were also undertaken to det. the elec. cond. of $PbCrO_4$ in agar agar soln. L. J. C.

Kinetic study of Liesegang rings. L. BULL AND SUZANNE VEIL. *Compt. rend.* 192, 682-3(1931), cf. *C A* 25, 1142 —The speed of formation of both primary and secondary rings varies as the square root of time. Primary rings form first but the secondary rings, forming more rapidly, finally extend beyond the primary. D. B.

Diffusivity of colloids—a method of determining the diffusion velocity of very slowly diffusing substances, based on a new measuring principle. II. R. BRUINS. *Rec. trav. chim.* 50, 121-8(1931) —An interferometer method was developed for measuring the speed of diffusion in colloidal solns., by which diffusion consts. 0.01 of those of the common electrolytes can be measured in 2 days. With a 0.7% starch sol, the results obtained were reproducible within 1 to 2%, and agree with the diffusion law. J. H. REEDY

The diffusion of colloidal particles. I. Abnormally high diffusion velocities in hydrophilic sols. II. R. BRUINS. *Kolloid-Z.* 54, 265-72(1931); cf. preceding abstr. —Diffusion velocities were detd. by the optical method at 20° for a wheat starch, unknown source, I, wheat starch, Merck *amylum solubile*, II, wheat starch, Huron Milling Co., III and gum arabic, 6% ash on dry basis. The velocities and radii of particles calcd. by the Einstein equation are: I, diffusion velocity 0.021×10^{-3} sq. cm./sec., radius 8.2 μ , II, 0.10×10^{-3} , 1.9; III, 0.15×10^{-3} , 1.4, gum arabic 0.25×10^{-3} , 0.85. The low calcd. radii are not compatible with the viscosity and slight diffusivity through membranes shown by the sols. Measurements on the lowering of the b. p. indicate that the

effect is not due to electrolytes. II. A new ion effect in hydrophilic sols. *Ind* 272 8.—The diffusion velocity of hydrophilic colloids decreases extraordinarily in the presence of low concns. of electrolytes, as shown for wheat starch III, and gum arabic. Expts in 0.1 N solns. of KCl, NaCl, and $\text{Co}(\text{NH}_4)_2\text{Cl}_2$ showed that the lowering increases with increase in valence of the oppositely charged ion. For a given ion the lowering decreases with increase in valence of the like charged ion. The neg. ion in this case. Gum arabic showed the following diffusion velocities times 10^6 : no electrolyte 0.225, 0.001 N KCl 0.078, 0.001 N K_2SO_4 0.002, 0.001 N $\text{K}_3\text{Fe}(\text{CN})_6$ 0.004 and 0.001 N $\text{K}_4\text{Fe}(\text{CN})_6$ 0.072. The relative lowering is greatest in the dil. electrolyte soln., but the selective effect of the ions is noticeable at 0.020 N solns. ARTHUR J. LEFICHER

Note on the Röntgen diagram of collagen. (Fiber period.) R. O. HERZOG AND W. JANCKE. *Z. Physik. Chem. Abt. B* 12, 228 9 (1931) cf. *C. A.* 21, 1213, K. Herrmann *et al.* *C. A.* 25 36.—The value for the fiber period, about 20 \AA , for collagen and elastin earlier obtained by H. and J. has been questioned by H. *et al.*, on the basis of their work on gelatin which indicates a value of about 9.5 \AA . H. and J. in Jha 121 the fact that at their preps. from stretched neck ligaments and heel tendons of cattle give much clearer diagrams than they could obtain from stretched gelatin. H. and J. obtained only 2 values of 0.107, 0.191 and 0.270. H. finds values agreeing with the first and the third but then considers only the first value in arriving at a fiber period of 9.5 \AA . EDEN J. SIFFARD

Swelling and syneresis of isoelectric gelatin jelly. E. J. BIGWOOD AND R. MAJNIV. *Compt. rend. soc. biol.* 104, 815 N. *Physiol. Abstracts* 15, 741 (1931).—The swelling of gelatin jelly depends on the total concn. of the protein in the setting liquid and the amt. of free and combined ions in the medium. The jelly consists of a solid reticular phase and a liquid interstitial phase. The elasticity of the former is opposed by the osmotic pressure of the latter. The salts influence the distribution of the protein in these 2 phases at the moment of setting, the protein passing more freely into the reticular phase at low ionic concn. and vice versa. In the latter case swelling of the gelatin occurs at lower concns. of gelatin. Curves of swelling made for blocks of jelly in acetic-acetate soln.—the same medium as the solvent of different strengths and at pH 4.5 and temp. 11° to 14° show that there is a crit. value of gelatin concn., for which there is no change of vol. in the blocks. A 5% jelly swells at N 3 concn. of ions, does not change at N 5% and shrinks at N 50%. G. G.

Swelling and hydration of gelatin. J. H. NORTON AND M. KUNITZ. *J. Phys. Chem.* 35, 102 4 (1931). A phase rule treatment of gelatin sols or gels adequately accounts for the various types of swelling characteristic of these sols. On the assumption that the gelatin-water system consists of 2 phases and 3 components, satisfactory explanations are obtained for the curves of osmotic pressure vs. concn. in g. per 100 cc. H_2O , the various viscous and hydration effects and finally the swelling and syneresis of isoelec. gelatin. A detailed picture of the gelatin micelles is deduced from the exp'tl. properties of the sols. H. R. MOORE

Hydrolysis of gelatin. B. N. DAS AND T. R. BOLAN. *Proc. 15th Indian Sci. Cong.* 1923, 144.—The effect of the hydrolysis of gelatin on its power to prevent the formation of Ag_2CrO_4 from solns. of AgNO_3 and K_2CrO_7 was studied. It was found that the inhibitive power of gelatin first decreases, then increases and on further hydrolysis again decreases. The pH of the gelatin solns. (as measured by the colorimetric method) is increased to a small extent by hydrolysis and shows little variation with the progress of the latter. By bringing the pH of hydrolyzed gelatin to that of unhydrolyzed by the addn. of dil. AcOH the inhibitive power of hydrolyzed gelatin is in no way increased, although the addn. of the same quantity of AcOH to the unhydrolyzed gelatin increases its inhibitive power a great deal. E. J. C.

X-ray study of the gelatinization of nitrocellulose. DESMAROUX AND M. MATHIEU. *Compt. rend.* 191, 74 8 (1930).—X-ray diagrams were obtained and described for films of nitrocellulose which contain various gelatins such as cyclohexanone, nitroglycerin and acetophenone. MALCOLM DOLE

Spectrochemistry of solutions of boric acid in glycerol. MORELLO MORELLI. *Atti accad. Lincei* 12, 451-C (1930).—The refractive indexes of solns. of H_3BO_3 in glycerol for the α & γ -lines of H_2 , as well as D line of Na, were detd. for concns. from 1 to 20% H_3BO_3 at 20° . The sp. refractivities $(n-1)^2/d$ and $\frac{(n^2-1)}{(n^2+2)d}$ were calcd., the sp. grs. of the solns. were also detd. The refractivity decreases with increasing concn., and is greater than in equiv. water solns. When H_3BO_3 is dissolved in glycerol there is an increase in vol. of the soln., amounting to 1.883 cc. at 20% concn. (vol. of 100 g. soln. calcd. 78.006 cc., measured 79.889 cc.) A. W. CONTIERI

Reducing action of sodium upon salts in liquid-ammonia solution. WAYLAND M. HURGRESS and EDWARD H. SMOKER. *Chem Reviews* 8, 265-72 (1931), cf. *C A* 24, 5218. Examples are given of the reduction of inorg. salts by Na in liquid NH_3 . Three types of reaction are illustrated: (1) the reduction product is the free metal, e. g., Ag salts are reduced to Ag by Na in liquid NH_3 ; (2) the reduction product is an intermetallic compound, e. g., Na reacts with $\text{Zn}(\text{CN})_2$ to form a black metallic ppt. which appears to be NaZn ; (3) the reduction product takes part in additional reactions, e. g., in the reaction of Na with MnI the free Mn formed not only catalyzes the reaction between Na and NH_3 but also appears to react with the NaNH_2 formed. L. K.

Properties of solutions of metals in liquid ammonia. WARREN C. JOHNSON and ALBERT W. MEYER. *Chem Reviews* 8, 273-301 (1931) cf. *C A* 24, 1270.—A discussion of (1) the solubility of the alkali metals in liquid NH_3 ; (2) mol. wt. detns. of metals in NH_3 ; and (3) the densities, activities, and elec. and photoelec. properties of such solns.

LOUISE KELLEY

Acid-base equilibria in non-aqueous solvents with particular reference to glacial acetic acid. NORRIS I. HALL. *Chem Reviews* 8, 191-212 (1931).—In acid-base equilibria solvents may be roughly classed as preponderantly acidic (e. g., AcOH), predominantly basic (e. g., NH_3 and amines), amphiprotic (e. g., H_2O) or aprotic (e. g., C_6H_6). Acid-base reactions in solvents of these 4 classes are discussed.

LOUISE KELLEY

The solubility of mercuric bromide in ethyl alcohol. KISHEN LAL and H. B. DUNNILLIFF. *Proc. 15th Indian Sci. Congr.* 1928, 168.—The values found by the method of Chuganov and Khlopov (*C A* 8, 1907) are consistently higher than those recorded by Timofeev. No further information is given.

F. J. C.

The solubilities of lanthanum oxalate and of lanthanum hydroxide in water. The mobility of the lanthanum ion at 25°. I. M. KOLTHOFF and RUTH L. MOUTAT. *J. Am. Chem. Soc.* 53, 1217-25 (1931).—Cond. measurements of LaCl_3 solns. gave a mobility of 77 at 25° for La^{3+} ion (Cl^- ion 75.8). A colorimetric method for the detn. of traces of La using Na alizarinate is described. The soly. of $\text{La}_2(\text{C}_2\text{O}_4)_3$ at 25° detd. volumetrically is 2.08 ± 0.03 mg. per l. The cond. method gave 0.0 mg., which is low because of dissociation of the dissolved oxalate. The soly. of the oxide and of the aged hydrous oxide was detd. at 0.7 mg. of La_2O_3 per l. at 25°. The cond. and volumetric methods agreed within 10%.

J. C. SNYDER

Solubility of acids in salt solutions. IV. The solubility of benzoic acid and the activity coefficient of its molecules in aqueous benzoate solutions. ERIK LARSSON. *Z. physik. Chem.* A, 153, 299-308 (1931) cf. *C A* 24, 4689.—The soly. of benzoic acid in solns. of LiCl , RbCl , CsCl , MgCl_2 , KBr , KI , LiNO_3 , NaNO_3 , KNO_3 , the chlorides and nitrates of Ca, Sr, Ba, NaClO , $\text{Na dichloroacetate}$, $\text{Na trichloroacetate}$, $\text{Na benzene sulfonate}$ and $\text{Na } \beta$ -naphthalenesulfonate were detd. The activity coeffs. of the benzoic acid mol. are calcd. V. Temperature coefficient of the activity coefficient of benzoic acid molecules in solutions of sodium chloride, potassium chloride and sodium benzoate. *Ibid.* 466-70.—The soly. of benzoic acid was detd. in KCl and NaCl solns. at 0.2° and 25°, in Na benzoate soln. at 25°. The activity coeffs. of the benzoic acid mol. were calcd. I. U.

The solubility and activity of silver benzoate and silver acetate in concentrated salt solutions. ERIK LARSSON and BIRGER ADELL. *Z. anorg. allgem. Chem.* 196, 354-63 (1931), cf. *C A* 21, 3292L.—The soly. of AgOz in solns. of Na, K and Ba nitrates and of NaOAc and the soly. of AgOAc in solns. of NaNO_3 and NaOAc were measured. In each case except AgOAc in NaOAc soln. the soly. of the Ag salt increases with increased concn. of the other salt. Calculs. of the av. activity coeff. of these Ag salts in salt solns. show that the dependence of the activity coeff. upon the 'ionality' of the

soln. can be computed by means of the Debye-Huckel formula, $-\log f = \frac{0.5\sqrt{\mu}}{1 + A\sqrt{\mu}} + B\mu$, by choosing values of suitable magnitude for the consts. A and B (f = the average of the activity coeff. of the 2 ions, μ = 'ionality').

W. C. FERNLIE

Determination of the molecular weight of cane sugar by the vapor-pressure method. S. OGURI, M. MATSUI, S. SHIMIZU and G. MONYA. *Rep. Tech. Chem. Lab., Waseda Univ.* 1928, No. 8, 6, *Mem. Faculty Sci. Eng. Waseda Univ.* 1930, No. 7, 13-6.—The mol. wt. of cane sugar was detd. by passing dry air through a soln. of known concn. and pure water, and noting the decrease in wt. of each.

L. W. T. CUMMINGS

The lowering of the vapor pressure of water by dissolved electrolytes. J. J. VAN LAAR. *Proc. Acad. Sci. Amsterdam* 33, 1140-60 (1930).—Existing data are shown to agree with the general formula $(p_0 - p)/p_0 x = 1 - \phi x^{1/2} + \phi x - \phi x^{3/2}$, where x denotes mol. fraction concn., ϕ van't Hoff's factor (e. g., $\phi = 2$ for a binary electrolyte) and ϕ a const. that can be calcd. from the theory of Debye and Huckel (e. g., approx. 5 for a

univalent binary electrolyte), while ρ and σ are individual consts. for the electrolyte in hand. Eventual assest. of the water has no effect on ρ and σ , it can only affect the consts. ρ etc.

The measurement of the conductance of electrolytes. IV. The validity of Ohm's law for electrolytes. GRINNELL JONES AND GILES M. BOLLINGER. *J. Am. Chem. Soc.* 53, 157-162 (1931), cf. *C. A.* 25, 1725.—If errors due to heating, polarization and the secondary effects of inductance and capacitance are avoided, there is no measurable variation in the real resistance of electrolytes with variation in the applied voltage. The applied voltage was varied from 0.01 to 5 v.

The mechanism of ion conductivity. ZOLTAN GYLLAT. *Mathemat. de Termis-ritud. Eszento* 47, 770-7. German abstract 778 (1930).—A soln. of PbCl_2 contg. 0.003 parts KCl does not follow the law of Ohm in an elec. field of 0-100 v./cm. elec. strength. The law of Ohm becomes valid in fields of larger elec. strength or in a soln. of pure PbCl_2 . Movable K^+ ions probably increase the cond. because of their contrary charge. They are thrown away in elec. charged fields and leave their former place. Addn. of BaCl_2 to PbCl_2 soln. did not cause such irregularity. Disturbance can be caused only by an ion whose contrary charge can loosen the ions that conduct the current.

The transference numbers of potassium chloride. New determinations by the Hittorf method and a comparison with results obtained by the moving-boundary method. DEAN A. MACINNIS AND MALCOLM DOLE. *J. Am. Chem. Soc.* 53, 1357-64 (1931).—Transference nos. for KCl were detd. at 25° for the concn. range 0.02 to 3 N by a slightly modified Hittorf method. The results agree with few values obtained by the moving-boundary method but they do not agree with old data obtained by either method.

The electrical conductivity at high temperatures of solutions of common salt and of concentrated sulfuric acid. I. J. SYMON. *J. Roy. Tech. Coll. (Glasgow)* 2, 395-401 (1931).—Measurement of the elec. cond. of aq. solns. of NaCl for concns. up to 1.13% and for temp. from 10° to 230° are described. Above 100° the solns. were necessarily under a pressure which rose to about 400 lb. per sq. in. at 230°. The cond. of H_2SO_4 from 94 to 99% was also measured for the temp. range 15-60°. Values given in the Smithsonian Tables (1920) for a few dil. solns. of NaCl may be in error by about 4% at 100° and 17% at 215°.

The influence of sucrose on the conductivity of electrolytes. J. PILLER. *Litz. Cukrarar* 49, 346-52 (1931).—The elec. conductivities of 0.01 N HCl , 0.01 N H_2SO_4 , 0.01 N KOH , 0.01 N Na_2CO_3 , 0.01 N NH_4Cl , 0.01 N CaCl_2 , and 10% AcOH were measured in the presence of 0, 5, 10 and 15% sucrose. Sucrose decreased the dissociation of the electrolytes and hence the cond., the decrease in elec. cond. was proportional to the sucrose present, ranged from 8 to 67%, and appeared smallest for the highly dissociated acids and alkalis. The order of decreased cond. was $\text{HCl} < \text{H}_2\text{SO}_4 < \text{CH}_3\text{COOH} < \text{Na}_2\text{CO}_3 < \text{KOH}$. Alkalis formed a compd. with sucrose and decreased the Cond. The depression of elec. cond. in salts increased with the valence. NH_4Cl , KCl , and NaCl were least affected (CaCl_2 and Na_2CO_3 more) by the presence of sucrose, the depressions of elec. cond. for the salts were greater than for the corresponding acid of equal concn. Salts of NH_4 were not affected in direct proportion to the sucrose addn. the addn. of an aq. sucrose soln. to NH_4OH increased the elec. cond., the max. cond. was reached with 16% sucrose, the original cond. was reached again with 36% sucrose and for further increased sucrose concns., the cond. continued to fall. AcOH and $(\text{NH}_4)_2\text{CO}_3$ decreased the elec. cond. for all sucrose concns. but not in a direct proportion to the sucrose concn. The decrease in elec. cond. for mixts. of salts with dil. sucrose showed $\text{KOH} > \text{KOH} + \text{KCl} > \text{KCl} > \text{KCl} + \text{HCl} > \text{HCl}$. In the presence of 50-80% sucrose the elec. cond. decreased in the order $\text{KOH} < \text{CaCl}_2 < \text{Na}_2\text{CO}_3 < \text{NH}_4\text{Cl} < \text{AcOH} < \text{HCl}$, dehydration by the sucrose is the cause for the reversal. The dielec. const. of sucrose (55) is high and decreased the dissociation of, particularly, the weaker acids, this decrease was only a portion of the total decrease in elec. cond. An increased viscosity, hydration of the ions, of the electrolytes, assocns. of sucrose mols., or of their hydrates, and a decreased ionic mobility are the causes of the remaining decrease. Chem. combinations may occur between sucrose and KOH , Na_2CO_3 , or NH_4 .

Aqueous solutions of sodium aluminate. I. Electrical conductivity. MATA PRASAD, S. M. MEHTA AND N. G. JOSHI. *J. Indian Chem. Soc.* 7, 973-80 (1930).—The Kohlrausch method was used to measure the elec. cond. of solns. contg. various ratios of Na_2O to Al_2O_3 . For dil. solns. no conclusion can be drawn as to the existence of aluminates. In concd. solns. a sharp change in equiv. cond. at a ratio of 3 to 1 is interpreted as indicating the formation of NaAlO_2 in soln.

Electrical conductivities of mercury amalgams of potassium and sodium. G. R.

PARANJPE AND BUIHARIWALA *Proc 15th Indian Sci Cong* 1928, 74.—The results (not given) of these detns indicate that in the elec cond curves the discontinuities do exist at the same places as those measured by Bhawe in the viscosity curve. This also corroborates the discontinuities measured by Bhatnagar and others as regards other phys properties. E J C

Neutral salt effect in ion reactions. IV. The specific ion effect. A. DE KISS AND I BOSSÁNYI *Acta Sci Univ Francisco Iosephinar, Acta Chem, Mineral, Physica* 1, 59-68 (1929) cf *C A* 24, 5382.—Neutral salt effect between persulfate and I ions was examd and data contrary to the theory of Brønsted were again obtained as found previously in coned solns. A theory is advanced in explanation. S S DE FINÁLY

The boundary layer of dilute electrolytes. II. MAX PLANCK *Sitzb preuss Akad Wiss* 1931, 113-22 cf *C A* 24, 4979.—The $e m f$ at a liquid junction should slowly change with time from that given by Henderson's formula to that of Planck. The former represents a mixt, the latter a stationary state. Since expt does not confirm this, another formula is presented which has a different concn gradient in the original boundary layer. This shows better agreement with expt. G M M

The electrocapillary curve of mercury. K BENEDWITZ AND K KÖCHLER. *Z. physik Chem, Abt A*, 153, 443-50 (1931).—The potential of an isolated dropping electrode may differ from that of a shorted electrode. Nevertheless, the existence of an inversion point at 0.5 v was again demonstrable. The addn of KNO_3 , KOH , KBr , KI , H_2S , KCN and KSH caused considerable displacement of the curve. A change in the direction of current flow occurred at the max. The latter does not necessarily indicate a true zero point. FRANK URRAN

The advances in the potentiometric determination of hydrogen ions. V ČURK. *Chem. Listy* 25, 79-84, 104-107 (1931).—The H , Sb , glass, O_2 and oxidation-reduction electrodes are reviewed. Bibliography of 60 references. FRANK MAKESH

Rapid electrometric determination of pH . J DI GLÉRIA *Metogasdandgi Kuletak* 2, 439-40 (1929).—The value of pH is read directly on a measuring wire with the aid of a regulable rheostat. S S DE FINÁLY

Hydrogen-ion concentration. G SCHAY *Kiserlet Koslemezsek* 31, 187-222 (1928).—A general description. S S DE FINÁLY

The use of the antimony-antimonous oxide electrode in the determination of the concentration of hydrogen ions and in potentiometric titrations. The Prideaux-Ward universal buffer mixture. HUBERT T. S BRITTON AND ROBERT A ROBINSON. *J. Chem Soc* 1931, 458-73.—The $Sb-Sb_2O_3$ electrode possesses a wide range of applicability in titrimetric work and can rapidly indicate pH values with a moderately high degree of accuracy. Electrodes made up with various types of Sb and Sb_2O_3 were studied. Those found most satisfactory consisted of a clean bar of Sb dipping into a soln contg suspended Sb_2O_3 . Vigorous mech stirring was necessary to obtain steady potentiometric readings. The electrode was calibrated as follows. To 100 cc of Prideaux-Ward universal buffer mixt (*C A* 18, 1002) was added 0.2 N $NaOH$ so that the soln passed continuously through a range of pH values from 2 to 12. The pH values were first obtained with a H electrode and with a quinhydrone electrode up to pH 9.1. A similar titration was then carried out with the Sb electrode and the $e m f$ s thus obtained were converted into the corresponding pH values. A N calomel electrode was used as the half cell and the junction was made with satd KCl . A series of titrations of acids and salts, some of which could not be used with a quinhydrone or H electrode, e g, HCN and H_2SO_4 , was carried out with the Sb electrode and the disson constns were detd where possible. In general, the pH values agreed well with those of other workers. The titration curve of H_2TeO_4 shows that it is a dibasic acid as opposed to the hexabasicity required by the formula H_6TeO_4 . With H_2SeO_3 and H_2SeO_4 the pH values were in good agreement with those of other workers, but the pH values could not be detd, probably because of a reaction of the Sb with the acid. The $e m f$ s obtained in the titration of H_2CrO_4 with alkali were not reproducible, in some cases the polarity of the cell being reversed. The Sb electrode was also found applicable in the titration of NaN_3 with HCl and the pH values were correct as far as can be judged from previous data. RUTH BERGGREN

Electrolytic reduction of dicarboxylic acids. HISAKAZU NAKATA *Anniversary Vol Dedicated to Masumi Chikashige (Kyoto Imp Univ.)* 1930, 49-55.—The electrolytic reduction of oxalic acid was studied in detail and the investigation was extended to malonic and succinic acids. Electrolysis of oxalic acid at 5-7° yields glyoxylic acid as well as substances more highly reduced. On prolonged electrolysis some glyoxylic acid undergoes further reduction, the yield decreasing from 89% in the second hr to 37% in the 6th hr. The presence of $HCHO$ could not be recognized in the cathode soln. The presence of

dihydroxytartaric acid could not be positively confirmed. At 60°, the current efficiency decreases but the yield of glycolic acid increases. The yield also increases directly with acid concn. The influence of the cathode material is very marked at higher temps. Hg is most suitable for obtaining glyoxylic acid, and Pb for glycolic acid. It was possible to raise the current and material yield to approx. 100% by decreasing the c/d to approx. 4 amp per sq dm with a Pb cathode. Under various conditions no reduction could be recognized in the electrolysis of malonic and succinic acids. ALLEN S SMITH

Interaction between hydrated copper oxide and neutral salt solutions. M. P. VANKATARAMA IYER. *Proc 15th Indian Sci Cong* 1928, 170-1.—Hydrated Cu oxide thoroughly purified by repeated washing and subsequent electroanalysis to remove the last traces of alkali is found by electromotive measurements to be pos. charged in contact with cond. water. It develops appreciable quantities of alkali when shaken with solns. of neutral salts. The pH developed depends upon the effect of the anion and not very much on the cation used. The variation of the elec. charge of the substance with various electrolytes can be correlated with the pH measurements. The total quantity of alkali liberated by repeated shaking of a given quantity of the sample with KCl and $BaCl_2$ was found to be a const. quantity. The pH , after attaining the value 7.6, remains const. on subsequent shaking with KCl. The sample of the hydrated Cu oxide was shown to absorb alkalis by measurements of the elec. cond. of the alkali soln. before and after adsorption. The results obtained strongly support the theory of the nature of the interface put forward by Mukherjee. I. J. C.

The reduction of manganese oxides by carbon monoxide. TISANUSO NISHIMOTO. *Anniversary Vol. Dedicated to Akazumi Chikashige* (Kyoto Imp. Univ.) 1930, 205-8.—The chem. equl. of the reduction of Mn_2O_3 and MnO by CO was investigated. The equl. const. of the reaction at 743° was detd. in various stages of reduction. In earlier stages of reduction the CO added was oxidized completely to CO_2 because of the high decompn. pressure of the Mn_2O_3 . In the stage of reduction before MnO the ratio $[CO_2]/[CO]$ suddenly decreased but the reduction could be followed beyond MnO . The latter yields on reduction a substance of strong reducing power capable of reducing H_2SO_4 to H_2S . ALLEN S SMITH

The mechanism of the oxidation of acetaldehyde and of hydrocarbons. MAX BODENSTEIN. *Z. physik. Chem.*, Abt. B, 12, 161-64 (1931), cf. *C. A.* 25, 2088.—The reaction scheme for the oxidation of hydrocarbons, previously discussed with reference to CaH_2 , is applied to the oxidation of C_2H_5 , (cf. *C. A.* 25, 1147). This is a more favorable example since the intermediate reactions lead to 3 different products. The expts. agree with the theory except in 2 cases where the concn. of the reacting gases was very small. The theory is also discussed with reference to explosion phenomena that take place in motors. G. M. MURPHY

Dissociation constants of nitrogen tetroxide and of nitrogen trioxide. FRANK H. VESHOER AND FARRINGTON DANIELS. *J. Am. Chem. Soc.* 53, 1270-63 (1931).—The dissocn. const. for gaseous N_2O_4 and the influence of pressure were measured at 25°, 35°, and 45°. The const. is proportional to the pressure and is 13% greater at 0.1 atm. than at 1 atm.; the const. is unaffected by O_2 or CO_2 . The dissocn. const. of N_2O_3 was measured at 25°, 35°, and 45° at different pressures. The influence of pressure on the equl. const. is due to deviation from the gas laws. The true equl. const. were detd. by extrapolation to zero pressure. ΔF , ΔH and ΔS are calcd. The dissocn. process according to $N_2O_4 \rightleftharpoons 2NO_2$ and $N_2O_3 \rightleftharpoons NO + NO_2$. S. LESHNIK

Heterogeneous equilibria between the sulfates and nitrates of sodium and magnesium and their aqueous solutions. MOHAMMAD A. HAMID AND AMBA PARSHAD. *Proc 15th Indian Sci Cong* 1928, 169.—The quaternary system $H_2O-Na_2SO_4-NaNO_3-MgSO_4-Mg(NO_3)_2$ has been investigated at 25°. In addn. to the solid phases in the ternary systems at this temp., a new surface appears in the quaternary system which probably represents the satn. field of the lower hydrate or hydrates of $MgSO_4$. I. J. C.

Role played by adsorbed gases in initiating reaction chains. combination of hydrogen and oxygen. HUBERT N. ALVAY. *J. Am. Chem. Soc.* 53, 1324-36 (1931).—The adsorption of H on glass increases above 450°. Adsorbed H initiates reaction chains extending into the gas phase at 600°. Cf. Alvay and Haber, *C. A.* 25, 635. S. L.

Are resonance phenomena possible in physical chemical periodicity? N. RASNEVSKY. *Z. Physik* 65, 270-2 (1930), cf. *C. A.* 23, 1041.—"Phys. chem. resonance" is suggested as a possible result of the interaction of 2 periodic reactions such as are assoc. with certain nerve phenomena. A math. treatment yields amplitude equations that are formally identical with corresponding equations applying to elec. and mech. systems. W. G. LEIGHON

Chemical activity and particle size: rate of solution of anhydrite below 70 microns. PAUL S. ROLLER, *J. Phys. Chem.* 35, 1133-42 (1931).—The soln. curves of 6 homogeneous size groups of cryst. anhydrite of surface mean diam. of 1.96 to 66.9 microns were detd. by cond. measurements. There is an induction period of 0.1 min. Below 25 microns the relative rate of soln. per unit surface increases to a max. of 17.6 at 2.8 microns and then decreases until at 1.96 microns it is 11.6. The soly. of natural cryst. anhydrite in H_2O is 0.298 g. $CaSO_4$, 100 cc. at 20° . S. LENHER

Statistical treatment of reaction-velocity data. II. L. J. REED AND E. J. THERIAULT, *J. Phys. Chem.* 35, 930-71 (1931). cf. *C. A.* 25, 2354.—A statistical treatment applicable to unimol. rates of reaction is developed. This procedure makes it unnecessary to det. experimentally the value of L in the unimol. formula, $K = [1/(t-t_0)] \ln[(L-1_0)/(L-1)]$. 1_0 and t_0 need not be accurately known. The usual methods of calcg. unimol. consts. is criticized. The least squares values of the velocity consts. of decomposition of N_2O_5 differ by 0.5 to 12.6% from values of Daniels and Johnston. The least-squares procedure should be used when a velocity const. is detd. within an allowable error of 5%. S. LENHER

Kinetics of transformation of the various forms and stages of hydration of calcium sulfate. WALTER FEITKNECHT, *Helv. Chim. Acta* 14, 83-90 (1931).—A comparison of Debye-Scherrer diagrams indicates that $CaSO_4 \cdot 2H_2O$ loses water with a change in lattice, to form a pseudomorph composed of disperse crystals of $2CaSO_4 \cdot H_2O$ between 10^{-4} and 10^{-5} cm. diam. (cf. Jung, *C. A.* 19, 1386). With respect to time, the dehydration of pptd. $CaSO_4 \cdot 2H_2O$ in *vacuo* passes through an induction period which at 20° , persists several hrs. without appreciable water loss, although the normal aq. tension is 12.7 mm. This suggests that dehydration sets in at isolated centers from which the process slowly spreads. The further dehydration of $2CaSO_4 \cdot H_2O$ to form sol. anhydrite proceeds, and may even be reversed, without disintegration of the crystals. There is an accompanying change in most of the weaker D-S lines. Partially dehydrated samples yield intermediate diagrams that do not show the two end lattices superimposed. It is concluded that the diffusion of water out of the semi-hydrate lattice causes contraction without changing the orientation. In effect the semi-hydrate is analogous to the zeolites, as pointed out by Linck and Jung (*C. A.* 18, 3566). The polymorphous transition from sol. to natural anhydrite is appreciable only above 300° , although sol. anhydrite is the unstable form. Here the transition consists of an abrupt change in the lattice, leading at first to a disperse product and followed by a growth in particle size at high temp. The latter fact explains the inertness of dead burned gypsum. W. G. LEIGHTON

The kinetics of the oxidation of copper. The establishment of sorption equilibria. F. J. WILKIN, *Phil. Mag.* 11, 422-32 (1931). cf. *C. A.* 24, 2365, 5208.—The quantitative treatment of the reaction kinetics of Cu oxidation by considering oxidation when the oxide O interface is not said, yet in equilibrium with the gas phase, leads to the equation $(p_s - p)^2 = [\alpha p_s (\ln(p_s/p) - 1) + \beta]/(p_s - p)^2 + \beta$, where p_s is the initial O pressure, p the pressure at time t and α and β are consts. For a series of expts. below the limiting pressure the curves are strictly linear after an initial deviation. The time at which linear behavior is reached increases with the no. of activations of the Cu. This is readily interpreted if activation increases the surface and it is assumed that adsorbed O is able to diffuse laterally. ARTHUR FLEISCHER

The measurement of hydrolysis in beryllium halide solutions. MILDA PRYTZ, *Z. anorg. allgem. Chem.* 197, 103-12 (1931). cf. *C. A.* 23, 5086.— $BeBr_2$ and BeI_2 were titrated electrometrically with NaOH at a series of water concns. and the data obtained were used to calc. the hydrolysis const. Although no true const. was found the best value was that which corresponds to $2Be^{++} + H_2O = Be_2O^{++} + 2H^+$. The assumption previously made that a monohydroxy ion is primarily formed which passes into a stable double compd. was confirmed. The α values of K_a found were 5.9×10^{-7} and 4.4×10^{-7} for bromide and iodide, resp. The soly. products for $Be(OH)_2$ were then calcd. as 2.7×10^{-19} and 3.3×10^{-19} as compared to 4.1×10^{-19} previously obtained from $BeCl_2$. R. H. LAMBERT

The hydrolysis of acetic anhydride. I. Hydrolysis of acetic anhydride in presence of neutral salts. J. E. KOCSIS, *Magyar Kém. Folyóirat* 37, 41-9 (1931).—Chlorides, nitrates, chlorates and bromides diminished but sulfates increased the velocity of hydrolysis. The salt influence law of Grube and Schmid (*C. A.* 20, 1548) can be applied to the hydrolysis of anhyd. $HOAc$. S. S. DE FINÁLY

Interaction of epichlorohydrin and cyclohexene oxide with alkali and ammonium halides. HEMENDRA K. SEN, CHITTARANJAN BARAI AND PATIT P. PAL, *Proc. 13th Indian Sci. Congr.* 1928, 146-7, cf. *C. A.* 21, 2668.—When these oxides act upon alkali and NH_4 halides a very interesting reaction takes place with the liberation of alkali

hydroxides and NH_3 . In fact this can be shown as a lecture expt. With an excess of NH_4Cl the velocity of reaction of epichlorohydrin (I) and cyclohexene oxide was detd. $1/\text{ratio } 1 \text{ } \text{NH}_4\text{Cl} : 1 \text{ mol } 10 \text{ mols}$ was used in the first expt. (0.925 g I, 5.35 g pure NH_4Cl and 20 cc pure abs alc were made up to 250 cc with distd water and placed in a thermostat at 35°). At 1 hr intervals 10 cc. of the soln. was taken out, at once mixed with an excess of cold water and titrated with 0.01 N H_2SO_4 . The velocity const. K after 1, 2, 3, 4 and 5 hrs. resp. = 0.2040, 0.1851, 0.1900, 0.2042, 0.2040. for the ratio 1 : 100 $K = 0.0423$ when the vol. of soln. is made up as described before, i. e., made up to 250 cc. If however the vol. is doubled and the same mol. proportion exists between the reactants $K = 0.4437$ when the vol. is quadrupled, $K = 0.2442$. Tables are given for the first 2 expts (run in duplicate). F. J. C.

Solubilities in the system water iodine to 200° . F. C. KRACKER. *J. Phys. Chem.* 35, 417-22 (1931). Water I solns above 112.3° form 2 liquid layers, the mutual soly increasing with temp. At the invariant temp the liquid layers contain 0.0317 and 99.3 mol % I resp. Below this temp the solns are satd. with solid I. The soly curves were detd. to temps above 200° by the use of the sealed tube method, by noting the temp. at which the last trace of the dissolving phase disappears. Invariant points were detd. by the method of thermal analysis. H. F. JOHNSTONE

Lowering of "ideal" eutectic temperature in n -component system. KEIZO IWASAKI AND NOBUYUKI NAKA. Anniversary Vol. Dedicated to Masumio Chikashige (Kyoto Imp. Univ.) 1930, 223. 7—A math. treatment based on the equation of an "ideal" m - p curve which holds in all the cases in which the solvent and solute are completely indifferent to one another. $T = T_0 / (1 - (RT_0/q_1) \log A_1)$, A_1 is the soly of the component 1, at temp. T . T_0 is its m . p., and q_1 is its heat of fusion. It is concluded that the eutectic point of n components is lower than the eutectic point of any $n-1$ components. ALLEN S. SMITH

Melting points and saturation points of sodium thiosulfate and sodium sulfate by the conductivity method. CHETAN ANAND AND HAR GORIND. *Proc. 15th Indian Sci. Cong.* 1928, 173.—The method was applied to $\text{Na}_2\text{S}_2\text{O}_3$ and Na_2SO_4 in aq. solns. These salts have comparatively low m . p's, are highly sol. in water and can easily be obtained in the supersatd. condition. Therefore if transition occurs at the m . p. in the solns., it will be indicated on the graphs. Resistances of various solns. were measured over a range of temps. which went above and below the m . p. of the salts and the satn. points of the solns. The resistance-temp. graphs for $\text{Na}_2\text{S}_2\text{O}_3$ show at its m . p. (33°) a clear transition point for all concns. and also a second transition point at the satn. point which varies with the strength of the soln. The graphs for Na_2SO_4 give two transition points. The m . p. varies between 43° and 50° with different concns. The graphs are not given. E. J. C.

Binary systems. III. J. H. KOUSS AND F. E. C. SCHEFFER. *Rec. trav. chim.* 50, 139-48 (1931). cf. C. A. 25, 628.—A math. treatment of crystn. of binary compts. from their dissoc. products, gaseous and liquid. J. H. REDDY

Thermal diagrams of the systems silver-strontium and silver-barium. F. WEINKE. *Z. anorg. allgem. Chem.* 193, 297-310 (1930).—A thermal analysis of the systems Ag-Sr and Ag-Ba was made. The following compts. with their m . p's. and d_{15}^{15} 's were found: Ag-Sr, 781° , 7.985-994; Ag₂Sr₃, 757° , 6.487-496; Ag-Sr, 680° , 5.695-711; Ag₂Sr₃, 666° , 5.151; Ag-Ba, 729° , 7.920-800; Ag₂Ba₃, 6493° -700; and Ag₂Ba₃, 648° , 6.597-603. Evidence of the existence of Ag₂Ba₃ was also found. The temps. of the eutectics were Ag-Ag₂Sr, 750° ; Ag₂Sr-Ag₂Sr₃, 693° ; Ag₂Sr-Ag-Sr, 628° ; Ag-Sr-Ag₂Sr₃, 645° ; Ag-Sr-Sr, 436° ; Ag-Ag₂Ba, 726° ; and Ag₂Ba-Ag₂Ba₃, 679° . The m . p. of pure Sr is given as $737^\circ \pm 1$. Exptl. difficulties did not permit the investigation of the Ag-Ba system for concns. of Ba above 60%. H. F. JOHNSTONE

X-ray study of the copper end of the copper-silver system. ROY W. DRIEK. *Ind. Eng. Chem.* 23, 404-5 (1931).—X ray examn. of various Cu samples indicates that Ag is not dissolved in the Cu. It is concluded that under ordinary room conditions Ag is not sol. in Cu. M. McMAHON

The system aluminum-silver. ENRICO CREPAL. *Atti III congresso naz. chim. pura applicata* 1930, 371-9.—Using the method of thermal analysis the equil. diagram of the system Al-Ag was redetd. Three compts. are formed (1) Ag₂Al at the peritectic temp. of 772° (the β form changes into the α -form at 609°), (2) Ag₃Al at the peritectic temp. of 727° and (3) Ag₄Al at the peritectic temp. of 693° . The compd. Ag₂Al forms a eutectic with a solid soln. of Ag in Al at 567° and a concn. of 26.5% Al (in the text 39.5 is given which must be a mistake since the table and the diagram clearly indicate

26.5%) The soly. of Ag in Al at the eutectic temp. is given as 45% Ag as derived from microscopic observation of specimens annealed at various temps. for different lengths of time. At room temp. the soly. is less than 0.5% Ag. These Al-rich alloys exhibit the phenomenon of age hardening. The soly. of Al in Ag at room temp. is 4.2%.

II S. VAN KLOOSTER

A thermodynamic study of the equilibria of the systems: antimony-bismuth and antimony-lead. CHU PHAY YAP. *Inst. Metals Division, Am. Inst. Mining Met. Eng.* (preprint) 1931, 24 pp.—For a long time the Sb-Bi system has appeared to contradict the phase rule in that it has an anomalous solidus always obtained by cooling down from the melt. Otani, by means of the resistance method, showed that it was a normal solidus characteristic of completely isomorphous systems. Y. applied the thermodynamic laws of the depression of f_p to the system in order to det. the cause of the anomalous solidus. The nonvariant crystn. is probably due to the formation of B_2 , although Sb and Bi should normally be considered diatomic. The heats of fusion of Sb and Bi are calcd. to be 20.0 and 14.2 cal. per gram, which agrees with the most reliable exptl. values obtained by Lumpe. The Pb-rich end of the Sb-Pb system also was analyzed thermodynamically. In this range Sb is dissolved in molten Pb in the monatomic form, although it is normally dissolved in solid soln. in Pb as Sb_2 down to about 150°. Below that temp., Sb is dissolved in the monatomic form. This suggests a transformation either of Sb or the Pb solid soln. at around 150° (probably less). The heat of fusion of Pb is calcd. to be 5.7 cal. per g., which is likewise in good agreement with reliable exptl. values. Attention is called to a new method of calcg. the heat of fusion of a solute.

C. L. MANTLE

Thermodynamic properties of dichlorodifluoromethane, a new refrigerant. II. Vapor pressure. W. K. GILKEY, FRANK W. GERARD AND MILO E. BUELER. *Ind. Eng. Chem.* 23, 354-6 (1931), cf. *C. A.* 25, 2340.—The vapor pressures of CCl_2F_2 from -70° to the crit. temp. (111.5°) were measured by a static method. An equation, suitable for thermodynamic calcns., is given: $\log_{10} p = 31.6315 - (1816.5/T) - 10.859 \log_{10} T + 0.007175 T$, where p is the pressure in abs. atms. and T is the temp. in deg. K. The calcd. normal b. p. is -29.8°C., the Trouton and Hildebrand consts. are, resp., 20.0 and 27.1. A table of calcd. vapor pressures at 10° intervals is given. III. Critical constants and orthobaric densities. F. R. BICHOWSKI AND W. K. GILKEY. *Ind.* 366-7.—The orthobaric ds. of CCl_2F_2 were measured up to the crit. temp., and the crit. consts. detd. The satd. vapor ds. below 50° were calcd. from the vapor-pressure equation and the equation of state. Above 50° they were obtained by detn. of dew points. Liquid ds. from -40° to 50° were detd. by a dilatometric method, and above 50° by Faraday's method, with glass floats. The crit. temp., pressure and d. are, resp., 111.5°, 39.56 atms., 0.555 g. per cc.

A. L. HENNE

The crystal structures of the compounds formed in the antimony-cadmium system. MASUMI CHIKASHIGE AND TOSHIOCHI YAMAMOTO. *Ann. Chem. Soc. Japan* 1930, 195-9.—An x-ray exam. to distinguish the stable and metastable compds. belonging to the Sb-Cd system was undertaken. The compd. $SbCd$ exists in two modifications, α and β which have the same space lattices with respect to type and dimension, belonging to the hexagonal system, the base and axial ratio of which are 4.42 Å. U. and 0.78, resp. The so-called metastable compd., Sb_2Cd , is a eutectic mixt. of β - $SbCd$ and Cd. Cf. *C. A.* 24, 1558. ALLEN S. SMITH

Crystalline structure of barium tungstate. II. J. PALACIOS AND I. NAVARRO. *Arteses soc. españ. fis. quim.* 29, 21-32 (1931), cf. *C. A.* 24, 1261.—The position of the O atoms in the elementary crystal prism of $BaWO_4$ was detd. From the position of Ba, W and ionic rays, approx. values of O coordinates can be deduced. The Moll microphotometer (*Proc. Phys. Soc. (London)* 33, 207 (1920-21)) was used. E. M. S.

Quantum mechanics of adsorption catalysis. M. BORN AND V. WEISSKOPF. *Z. physik. Chem., Abt. B*, 12, 206-27 (1931), cf. *C. A.* 25, 21.—By a perturbation calcn., the theory of Born and Franck is elaborated to include interaction between the adsorbed mol. and the adsorbing crystal surface. By applying the method of Weisskopf and Wigner (*C. A.* 24, 5602), the order of magnitude of the reaction rate is found as a function of an at. distance and a vibrational quantum no. An appendix discusses some of the math. difficulties connected with the calcn. of the probability amplitudes and the matrix elements.

G. M. MURPHY

Catalytic decomposition of nitrous oxide. M. S. SHAH. *Proc. 15th Indian Sci. Cong.* 1928, 170.—In the analytical investigation of the reaction $C + N_2O$ at various temps. it was found that N_2O decomps. at a lower temp. in the presence of charcoal than when heated alone. This suggests that the behavior of charcoal is catalytic. On examn. of the catalytic influence of ThO_2 , Al_2O_3 , TiO_2 and Pt a similar lowering in the

temp of decompn of N_2O was observed. Comparison of curves for N_2O decompd in an hr against temp in the presence of these substances showed that these substances act as catalysts in the order ThO_2 , charcoal, Al_2O_3 , Pt black, TiO_2 , and Pt foil. I. J. C.

Reactions at the surface of hot metallic filaments. V. Thermionic emission and catalytic activity at the surface of hot metallic wires: $H_2 + CO_2 \rightarrow CO + H_2O$ at the surface of platinum, platinum coated with barium oxide and thoriated tungsten. B. S. SRIKANTAN. *Indian J. Physics* 5, 687-94 (1930). cf. C. A. 25, 1149.—The min temp at which thermionic emission from a Pt wire is perceptible was obtained by extrapolation as 975°. Interaction of CO_2 and H_2 starts at 979°. Hence, emission of electrons from a surface has an important influence on activation of gases. L. P. HALL.

Reactions of hydrogen and oxygen on platinum wires at low temperatures and pressures. HERBERT G. TANNER AND GUY B. TAYLOR. *J. Am. Chem. Soc.* 53, 1280-95 (1931).—A study of the catalytic combination of H_2 and O_2 at low pressures on Pt wires shows that the reaction is much more complex than ordinarily supposed. The consumption of excess O_2 from 2:1 H_2 : O_2 gas mixts is attributed to the formation of H_2O_2 . The rate of reaction varies greatly as a function of the temp difference between the catalyst and the wall of the contg vessel. Flashing the Pt wire catalyst in O_2 , H_2 or vacuum at 900° induces a temporary superactivation capable of causing reaction at a temp as low as -120°. P. H. ILMETT.

Decomposition of diethyl ether in contact with platinum and tungsten. H. AUSTIN TAYLOR AND M. SCHWARTZ. *J. Phys. Chem.* 35, 1041-53 (1931).—Decompn of $(C_2H_5)_2O$ in the presence of heated filaments of Pt and W is identical with the homogeneous reaction. The presence of the filament merely furnishes the necessary stationary concn of active mols. S. L. FETTER.

Esterification in the gaseous phase with solid catalysts. N. G. GAJENDRAGAN. *Proc. 15th Indian Sci. Congr.* 1928, 148.—The esterification of MeOH and EtOH with AcOH was studied with γ -alum and silica catalysts at 230°. The equil const. (not given) was not far from that obtained by other observers in the liquid phase at lower temps. I. J. C.

Active aluminum. G. SAMBAMURTI AND N. L. NARASIMHAN. *Proc. 15th Indian Sci. Congr.* 1928, 168.—Al becomes activated by amalgamation and the activated metal serves as a good reducing agent in a neutral medium. The activated metal readily oxidizes in moist air having traces of CO_2 . I. J. C.

Catalytic reaction between stannic oxide and lime. SETSURO TAMARU AND NORIKU ANDO. *J. Chem. Soc. Japan* 52, 36-46 (1931), cf. C. A. 25, 2045.—Cassiterite and artificial SnO_2 are acted on by acids and bases with great difficulty, e. g., they are made sol by melting with NaOH, Na_2CO_3 , $NaHSO_4$, etc. While no change occurs on heating them with lime at about 900-1000° in the presence of some catalytic substances they react with lime easily at these temps, being transformed into an acid sol form. H_2 , C, CO, S, Zn, SnO and traces of org. vapors with reducing power are very effective as catalysts. It is interesting that a reaction between solids proceeds smoothly at a temp below the m. p. of each component. The optimal conditions are the proportion 1 mol SnO_2 + 7 mol CaO + 0.05 mol H_2 or C, reacting temp 900°, heating for 1 hr., and the presence of H_2O . From the fact that the curves of catalytic action of H_2 and C resemble those of reduction, it is inferred that reduction takes place as the first stage of catalysis. With catalysts such as C which produce no H_2O as a reduction product, almost no reaction occurs without addn of H_2O , but where H_2O is produced the reaction proceeds smoothly. The quantity reacting is proportional to the logarithm of the duration of heating and to the no. of mols of CaO , the latter effect depends in reality on the surface area of CaO . T. IRIB.

The entropy and free energy of methane. H. H. STOKES. *J. Am. Chem. Soc.* 53, 1266-9 (1931).—The entropy of CH_4 at 25° obtained experimentally checks the value 43.39 entropy units previously calcd from sp. heat data. B. H. LAMBERT.

The heats of neutralization of eugenol and isoeugenol. G. GUNDU RAO. *Proc. 15th Indian Sci. Congr.* 1928, 147.—The heats of neutralization of these 2 isomers were detd to study their acidic behavior. The value for eugenol at about 25° when a slight excess of alkali was used was found to be 6476 cal. per mol and 6790 cal. per mol with a large excess of alkali. Because of incomplete soln when an equiv. amt of alkali was employed and partial pptn of the salt with a large excess, the value for isoeugenol, 6550 cal. per mol., is only approx. F. J. C.

Revision of the free energy of formation of sulfur dioxide. E. D. EASTMAN. *Bur. Mines Information Circ.* 6454, 6 pp (1931).—Recalculs described lead to the following revised equations: $2S$ (rhomb) = S_2 (g) $\Delta F_{298}^\circ = 18,280$ $\Delta F^\circ = 29,840 - 0.067 \ln T + 4.55 \times 10^{-4} T^2 - 39.81 T^{-1/2}$; S_2 (g) + O_2 = SO_2 (g) $\Delta F_{298}^\circ = -79,590$ cal.,

$\Delta F^\circ = -84,905 + 3.21 T \ln T - 2.272 \times 10^{-4} T^2 + 0.16 \times 10^{-6} T^3 + 0.25 T$. S (rhomb.) + $O_2 = SO_2$ (c) $\Delta F^\circ_{112} = -70,440$ cal ALDEN II 1 MFRY

The specific heats of sodium sulfate solutions. MAURICE LUMÉRAS. *Compt rend.* 192, 359-61 (1931) The sp heats of aq solns of Na_2SO_4 were detd at 20° at concns ranging from satn to extreme diln. The detns were made by means of a calibrated thermophore which was immersed in the solns. and gave check re values on successive detns. The sp heat compn curve obtained in this manner is somewhat higher than that obtained from the data of other investigators but the two curves approach each other at higher dilns. I W LAIRD

Calculation of the heat of distillation. M BARTHEL. *J Phys radium* [7], 1, 411-5 (1930) By math analysis an expression is obtained for the flow of heat for vapor and binary liquid in which only the compns of the vapors entering and leaving are considered. L P HALL

The formation of films of drying oils (KAPPELMAYER) 26. Fractional macro- and microsublimation at ordinary and at reduced pressures (ILLARI) 1. Metallic ions as catalysts for the removal of SO_2 from boiler furnace gases (JOHNSONS) 21. Telephone as zero instrument in the electrometric compensating method (HAROS) 7. State of CH_3O in aqueous solutions (WALKER) 10.

Journal of Science of the Hiroshima University Series A (Mathematics, Physics, Chemistry) (New journal) Published by the Hiroshima University, Hiroshima, Japan Vol 1, No 1 appeared December 1930

Annual Survey of American Chemistry, 1930 Vol. V. Prepared by the Division of Chemistry and Chemical Technology of the National Research Council. Edited by C J WEST New York The Chem Catalog Co, Inc 160 pp \$5.

BRIGGS, T C., AND TILMAN, H H Master Minds of Modern Science. London G G Harrap & Co, Ltd 278 pp

BRIGGS, DENNIS D Chemical Change. Toronto J M Dent and Sons, Ltd 160 pp 5s

FRITZSCHE, WALTER Über topochemische Umsetzungen fester Stoffe in Flüssigkeiten. Berlin Gehringer Hornträger 50 pp M 7

FLEURY, P Couleurs (étude physique) et colorimétrie: Conférence faite au Conservatoire National des Arts et Métiers le 7 mai 1930 Paris Hermann & Cie. 34 pp 1. 5

GREGORY, JOSEPH C. A Short History of Atomism: from Democritus to Bohr. London A and C Black, Ltd 238 pp 10s 6d, net

HOLMARD E J Practical Chemistry for Beginners. Toronto J M Dent and Sons, Ltd 142 pp 2s

HUME ROTHLEY W The Metallic State: Electrical Properties and Theories. London Oxford Univ Press 372 pp 25s, net

KOLTHOFF, I M The Colorimetric and Potentiometric Determination of pH . New York John Wiley & Sons, Inc

LAFAY, A Cours de physique. Tome II. Thermodynamique, optique. Paris Guithier-Villars et Cie 736 pp. F 150 Cf C A 25, 1150

LIEPMANN, EDMUND O VON Entstehung und Ausbreitung der Alchemie. Band II. Berlin J Springer 258 pp M 24, bound M. 26 60 Cf C A 14, 686.

LUDWIG, HERBERT Table for Determining Porosity in Per Cent by Volume. Mannheim Friedrichsfeld (Baden) the author. M. 1. Reviewed in *Sprachsal* 64, 57, *Ceram. Abstracts* 10, 592 (1931)

MAASS, O., AND STRACIE, E W R An Introduction to the Principles of Physical Chemistry. New York John Wiley & Sons, Inc

MRSICE, ROY O Work Book of General Chemistry. 3rd ed. Kalamazoo, Mich Kalamazoo Board of Education 99 pp

MITCHELL, C. AINSWORTH The Scientific Detective and the Expert Witness. Revised and enlarged ed of "The Expert Witness." London Simpkin Marshall, Ltd 242 pp 3s 6d, net

NORTHROP, F S C. Science and First Principles. New York The Macmillan Co. 200 pp

THEO, FRANK L. Torricelli contra Mundum. London. H. K. Lewis. 82 pp 5s. Reviewed in *Chem News* 142, 190 (1931).

3—SUBATOMIC PHENOMENA AND RADIOCHEMISTRY

W. ALBERT MOYES, JR.

Investigations on the structure of matter. B. LAMM. *Technika* 11, 69-74 (1930) —
A general summary. S. S. DE FINALLY

A rotating light-quantum (photon) model. HANS BAUER. *Physik Z.* 31, 882-8 (1930).
LOUIS WALDMAN

The nature of the chemical bond. Application of results obtained from the quantum mechanics and from a theory of paramagnetic susceptibility to the structure of molecules. LUDWIG PAULING. *J. Am. Chem. Soc.* 53, 1367-1400 (1931) — The electron-pair bond is discussed and from quantum mechanics a set of rules is presented which describes the properties of the bond with special reference to the strength of the bond and the nature of the single electron proper functions. These rules give information about the relative strengths of bonds formed by different atoms, the angles between bonds, properties of tetrahedral atoms with single and double bonds, *cis* and *trans* isomers, the *sp* and spatial configuration of bonds and other properties. Transitions from electron pair to ionic bonds are also discussed. A theory of the magnetic moments of mol. and complex ions is also developed. For the transition elements the proper functions involved in bond formation show that compounds with CN have electron pair bonds, those with F have ionic bonds, and those with H_2O , van dipole bonds. Electron structure, bond angles and other properties of mol. and complex ions can also be deduced from the magnetic data. C. M. MURPHY

Avogadro's number and electronic charge. SATYENDRA RAY. *Proc. 15th Indian Sci. Cong.* 1928, 60. E. J. C.

The rotating electron. M. N. SARA. *Proc. 15th Indian Sci. Cong.* 1928, 62. E. J. C.

The imaginary character of the "mass equivalence of electricity." SATYENDRA RAY. *Proc. 15th Indian Sci. Cong.* 1928, 68. E. J. C.

The conception of the Hall effect. H. ZANN. *Naturwissenschaften* 18, 848-9 (1930) — Z. concludes that it is necessary to distinguish between an "electron gas" Hall effect and a structural Hall effect. R. H. FRUTKINSON

Mean free path of electrons. Answer to a notice by C. Ramsauer. J. S. TOWNSEND. *Ann. Physik* 8, 805-8 (1931) — T. defends his indirect measurements against the criticism of Ramsauer. F. J. ROSENBAUM

Method of measuring the effective cross section. Answer to J. S. Townsend. CARL RAMSAUER. *Ann. Physik* 8, 109-10 (1931), cf. preceding abstr. — R. repeats his criticism of Townsend's measurements. F. J. ROSENBAUM

The electron-beam discharge in a glow. J. DOVEY. *Proc. Leeds Phil. Soc.*, Sci. Sect., 2, 198-8 (1931) — The collapse of the electron beam to a glow surrounding the anode as previously described (cf. Whiddington, *C. A.* 24, 2044) is due to the negative charging of the walls of the cone. When the tube is lined with Cu gauze the beam has a velocity corresponding to the voltage applied. The potential distribution in the discharge was investigated with an exploring electrode. With a thermionic current of 0.2 ma., a pressure of 0.002 mm., an anode potential of 242 v. and a length of glow of 0.85 cm. the following results were obtained (the first figure is the distance from the anode in cm., the next is the space potential in v.): 0, 242, 1.09, 60, 1.04, 54, 1.72, 41.5, 2.31, 34, gauge. 2. The edge of the glow is maintained at 60 v. and has a neg. space charge. J. B. AUSTIN

The potential of the walls in the cathode dark space. JEAN W. DECK AND E. C. FENELÉAS. *Phil. Mag.* 11, 54-64 (1931) — The walls and discharge are at the same potential at one place, i. e., the boundary between the neg. glow and the cathode dark space. The effect of the walls in Brown and Thomson's (cf. *Phil. Mag.* [7], 8, 918 (1929)) work is not of greater uncertainty than the incomplete knowledge of the effect of disson of gases or of the interpretation of the counts of Morse's equation A. F.

The capture of electrons by protons. KARL WOLF. *Ann. Physik* 7, 337-48 (1930), cf. Mark and Wolf, *C. A.* 24, 5500 — The procedure previously described was modified by the adoption of a photometric electrometric method of measurement. No selective absorption of electrons by the proton stream could be detected under energy relations predicted by analogy to the conditions described in case of a particles, by Davis and Barnes (*C. A.* 24, 1029) and Barnes (*C. A.* 24, 3264). W. G. LEIGHTON

Capture of electrons from mercury atoms by positive ions of helium. C. F. POWELL AND A. M. TOWNSEND. *Nature* 127, 592-3 (1931), cf. *C. A.* 25, 456 — Expts. in a new app. in He at 20 mm. pressure show a peak in the current frequency curve for each

type of ion present. The higher frequency peak corresponds to the He ions, while a 2nd, lower frequency peak is observed when a trace of Hg is present, indicating the formation of Hg^{++} formed by electron capture from He^+ . The ratio of mobility of Hg^{++} in He to He^+ in He is calculated to be 0.53. G. CALINGAERT.

Experiments on the efficiency of an electron gun. J. E. TAYLOR. *Proc. Leeds Phil. Lit. Soc., Sci. Sect.*, 2, 169-73 (1931).—An improved type of electron gun is described in which better focusing and higher beam efficiency are obtained by interposing between the filament and muzzle a neg. screen with a small central hole. This screen has low neg. potentials applied to it. The device resembles the soft 3-electrode valve in its action. J. B. ARSTIN.

The distribution of space-potential in high-frequency glow discharge. D. BAYERJ AND R. GANGLI. *Phil. Mag.* 11, 410-22 (1931). *C. A.* 24, 5612.—In rarefied air and in O_2 with a current at a frequency of 10 cycles, the potential difference is sym. about the middle point of the discharge tube. The extension of the glow outside the electrodes is also a sym. reproduction of the phenomena inside, with the difference that after a certain distance the potential is reduced to 0. ARTHUR FLEISCHER.

The emissivity of liquid iron alloys. GERHARD NAESER. *Mit. Kaiser-Wilhelm-Inst. Eisenerforsch. Dusseldorf* 12, 365-72 (1930).—The radiation temps. of metals depart considerably from the true temps. A correction of the temps. detd. optically is therefore always necessary, but, because of the uncertainty of the emissivity, the correction is not reliable. Expts. were made by the author to redet. the emissivity of liquid Fe-C alloys, of several alloy steels and of several slags. R. RIMBACH.

Positive-ion emission from thin platinum films on glass. R. A. NELSON. *Rev. Sci. Instruments* 2, 173-9 (1931).—In continuing expts. with Pt plated glass filaments (*C. A.* 25, 877) variations in pos. ion current were investigated. A uniform electrolysis potential was applied to all parts of the emitting surface of the app. Curves were obtained showing the variation of pos. thermionic current with electrolysis potential for K and Na glass, the existence of pronounced max. and min. with these equipotential emitters is apparent. An explanation of the phenomena is offered based on analogous electron emission from oxide-coated filaments and Cs-coated W. M. McMAHON.

A method of examination of transient glows and their spectra. H. PARAMESWARAN. *Proc. 15th Indian Sci. Congr.* 1928, 78.—This is a method suitable for the visual observation, photographic and spectroscopic studies of transient glows observed in some rarefied gases under chem. excitation during and after discharge. The device used consists of a Hg break (mounted on a base that can be rotated by a slow motion arrangement) carrying on its extended spindle a disk with an adjustable aperture in it. The glow is observed through this aperture. E. J. C.

Investigations of the inertia of gas-filled photoelectric cells. FRITZ SCHROTER AND GUNTHER LUBCZYNSKI. *Physik. Z.* 31, 897-904 (1930).—Photoelec. cells, contg. Ne or He and KH, show a certain current after the exciting light has been cut off, due to the ionization by collisions of the 2nd kind of the metastable ions. In metastable A, the energy is insufficient to cause further ionization, while in H_2 , there are no metastable states. Expts. were made with pure A and KII and with A, KH, and traces of H_2 . Cells of this type gave identical results when used for television, and gave no lag. Cells employing K satd. with H_2 at operating temp. gave satisfactory television results. LOUIS WALDBAUER.

Photoelectric effect on thin layers of potassium and cesium. R. FLEISCHER. *Physik. Z.* 32, 217-8 (1931).—By using a high-vacuum arrangement, a max. yield of 37.9×10^{-12} coul./cal. was obtained for K at 436 m μ . Cs gave a max. yield of 17.1×10^{-12} at 492 m μ . The yields were, resp., 25.7% for K and 10.3% for Cs. C. E. P. J.

Relation concerning atomic nuclei. WM. D. HARKINS. *J. Am. Chem. Soc.* 53, 2009-11 (1931).—Comments on Latimer's paper (*C. A.* 25, 2359). C. J. WEST. Structure of the α -particle. O. K. RICE. *J. Am. Chem. Soc.* 53, 2011-2 (1931).—Comments on Latimer's paper (*C. A.* 25, 2359). C. J. WEST.

Origin of cosmic radiation. J. H. JEANS. *Nature* 127, 504 (1931).—In calcg. the frequency ν of cosmic radiation from its absorption μ , nuclear as well as extranuclear electrons should be taken into consideration. This doubles or more than doubles the absorbing power of all atoms except H. A table showing the absorption (per meter of H_2O), for the radiation produced by various processes is given. Agreement between calcd. and observed values is excellent. This suggests that the most penetrating constituent may originate in the annihilation of an α -particle and its 2 neutralizing electrons, and the next softer one from 1 proton and its one neutralizing electron. G. C.

The age of the earth—radioactivity methods of its determination. ALORS F.

AuCl₃ is volatilized in the tube without removing the deposit of Re₂O₇, the presence of Cl brings out the Re spectrum, although no Au lines are visible. The Re doublets are repeated 16-32 and 48 units higher as ReO, ReO₂ and ReO₃. Re consists of 2 isotopes, 185 and 187, estd photometrically to exist in the ratio 1:1.62. Its packing fraction, estd from line 203 of the Re¹⁸⁷O is -1 ± 2 , the same as that of Os. The calcd at wt is 186.22 ± 0.07 . G. CALINGAERT

Methods of investigating the intensities of γ -rays. C. D. ELLIS AND D. SKOBELTZYN. *Nature* 127, 125(1931).—The Skobeltzyn method (C. A. 24, 1573) of investigating the intensities of γ rays depends on observing by the Wilson cloud method the relative no. of electrons ejected within a certain angular range by the Compton effect of the different γ rays. The Ellis and Aston method depends upon the photoelectric effect. γ rays of Ra B and C showed intensities of 0.22, 0.46 and 0.26 by the photoelectric method after passage through 3.5 mm Pb. The corresponding intensities by the Compton effect were 1.18 and 1.19, resp. This agreement is as good as can be expected. FRANK MARSH

The question of the disturbance of the equilibrium between radium B and radium C in preparations freed from residual emanation. KARL MARRACH. *Sitzb. Akad. Wiss. Wien, Abt. IIa*, 139, 211 9(1930).—The usual methods of removing emanation from preps. have been questioned as to whether or not they disturbed the Ra B-Ra C ratio (cf Braddick and Cave, C. A. 23, 1048). Removal of emanation by the usual short time washing with alc. heating to 400° or blasting with a strong current of air is believed rarely to disturb the Ra B-Ra C ratio as much as 1%. O. E. S.

Scattering of hard γ -rays. II. L. H. GRAY. *Proc. Roy. Soc. (London)* A130, 524-41(1931), cf C. A. 24, 5220.—The relative scattering powers of Mg, Al, Cu, Zn, Cd, Sn and Pb for Ra C γ rays filtered through 4 cm of Pb and for Th C γ rays filtered through 3 cm of Pb have been investigated by observing the ionization produced by radiation scattered within the angular range 10-30°. For Ra C γ -rays the decrease of scattering power per extranuclear electron from Mg to Pb is 2%, for Th C γ rays there is a gradual increase of about the same amt, but this is of the order of magnitude of error in the application of the correction for reabsorption of scattered radiation in the scatterer. Therefore these measurements establish the fact that the binding forces in heavy elements do not influence appreciably the scattering power of extranuclear electrons for radiations of energy corresponding to 2×10^4 electron v. J. E. A.

Radioactive metamorphic phenomena in fluorspar from Wölsendorf. LUIS GOEBEL. *Sitzb. Akad. Wiss. Wien, Abt. I*, 139, 373 9(1930), cf Mägge, C. A. 17, 3644.—The radioactive "halo" of radioactive fluorspar from Wölsendorf are nothing other than aged ring halos as shown by development series studies and range measurements of a particles. The colloidal nature of the pigments in Wölsendorf fluorspar is shown by use of the ultramicroscope. The colloidal particles are most frequently Ca particles and the color changes according to the size of the particles, small particles correspond to green, larger to blue, violet, etc. The radiations from the radioactive materials probably cause the formation of neutral Ca and F from the fluorspar as "stinking fluorspar" (contg free F and usually dark colored) is rather common. O. E. S.

Nature of the groupings of radioactive atoms. E. L. HARRINGTON. *Compt. rend.* 192, 414-5(1931).—The conclusion of Jedrejowski that the formation of aggregates of radioactive atoms is facilitated by impurities, groups forming around the nuclei of impurities, has been confirmed. But it is found that the impurities are not essential. The absence of aggregates on freshly cut mica is explained in a different way. The aggregates are more numerous around scratches, detached scales and along edges. Either the groups form at the position observed or they form in the atm. Polar mols. also aid in the formation of groups. The grouping of radioactive atoms which form in mixed gases contg Rn are constituted of the radioactive atoms and polar mols. around a nucleus. It is usually a stable assocn of ionized radioactive atoms. The beds of polar mols. have a low vapor pressure. The detection of aggregates depends on the no. of radioactive atoms as well as the no. of polar mols. L. D. ROBERTS

The scintillation of calcium tungstate. BERTA KARLIK. *Sitzb. Akad. Wiss. Wien, Abt. IIa*, 139, 319-26(1930).—The scintillation of a no. of specimens of artificial Ca tungstate under α and γ -rays was investigated at various pressures. With α -rays the brightness of the scintillated light as measured by a photoelectric cell is proportional to the residual range of the α -rays but reaches a satn. value for very small bodies. The decompn. with γ -rays is of the same order as that of ZnS but is unusually small with α rays giving Ca tungstate a unique position as a scintillating substance. E. A. FLOOD

Quantitative investigations on the photographic action of α - and H-particles.

MARIETTA BLAU. *Sil. Akad Wiss Wien, Abt. IIa*, 139, 327-47 (1930) — From the tracks produced by α and H particles striking a photographic plate at an angle ($20-24^\circ$) the no. of particles and distribution ranges were detd. and found to agree with the values obtained with a tube electrometer. The α particles were from Th C and the H particles from paraⁿ. "Agfa x-ray dental film" and "Imperial process" plates were used. The range distribution showed a displacement toward smaller values, due to scattering in the emulsion. The no. of H particles obtained on using the dental film was only half that obtained by the "Imperial process" plate or by the electrometer. The apparent distance between the AgBr grains of the dental film for H rays was twice as great as for α rays. F. A. FLOOD

The mechanism of separation of radon from radioactive minerals in liquid medium. V. I. SPRIZIN. *Tr. radium et minerais radioactifs acad. sci. U. R. S. S.* 2, 264-71, cf. C. A. 23, 1048 — The basic purpose of this investigation was to find general laws which govern the process of transition of Rn from minerals into liquid media so that with their help the variation in radioactivity of various springs could be explained. As far as the qualitative side is concerned, the results obtained permit the making of satisfactory conclusions. Thus only tonarite, a mineral which is found comparatively rarely, has no pos. temp. coeff., and therefore it could be said with a sufficient degree of probability that hot springs will be placed in more favorable conditions as to the absorption of Rn than cold ones. This conclusion could not be obtained by the observation of springs proper because several factors act simultaneously which combine to complicate the situation. No results were obtained on chem. compn. of mineral water and radioactivity. It is shown that with typical, purely uranium minerals the concn. of NaCl soln. increases considerably the rate of sepn. of Rn. Samarskite and tonarite liberate in NaCl soln. the same amt. of Rn as in water. In no case is there a considerable decrease in the rate of sepn. of Rn. A. C. NOB

The formation of molecular aggregates in radon-gas mixtures containing polar molecules. E. L. HARRINGTON AND O. A. GRATIAS. *Phil. Mag.* 11, 285-96 (1931); cf. C. A. 23, 230 — The formation of aggregates of active deposit atoms in gases containing radon in the presence of HCl , SO_2 , EtOAc , *o*-nitrotoluene, *p*-nitrotoluene, CHCl_3 , CCl_4 , MeOAc and acetone was studied by the centrifuge, the photographic and the ultra-microscopic methods. The previous work on air radon H_2O and water free air radon mixts. was confirmed. Aggregate formation in the presence of HCl is of av. magnitude; it is very large in the presence of *o*-nitrotoluene. Polar mols. favor the formation of aggregates in radon gas mixts. ARTHUR FLEISCHER

The absorption of scattered x-rays and Compton's theory of scattering. S. R. KHASTGIR. *Proc. 15th Indian Sci. Cong.* 1928, 83-9 — The most significant feature in the comparison of the absorptibilities of the secondary (scattered) and primary x-rays by the method of Barkla and K. (C. A. 20, 146) is the constancy, over a wide range of wave lengths, of the ratio of the ionizations produced by the secondary and primary beams, when the beams are either *unintercepted* or *intercepted* by similar sheets of an absorbing substance. The 2 alternative cases recorded are: A (not considered here) in which discontinuities in the "intercepted" ratio S'/P' are obtained and B, the only case giving results at all compatible of explanation in terms of Compton's theory of wave-length change on scattering, in which the values for the "unintercepted" ratio S/P and the "intercepted" ratio S'/P' for various wave lengths are found to lie on 2 parallel horizontal lines without the slightest sign of any discontinuity. Calculations of the ratio S'/P' were based on C. T. Ulrey's energy-distribution curves (C. A. 12, 1530). From these curves of U, corresponding curves were then constructed for both primary and scattered beams intercepted by an equal thickness of an absorbing material. In the case of the scattered beam, the increased value of the absorption coeff., corresponding to Compton's wave length change was used. The presence of an unmodified radiation in the scattered beam was also taken into account. Thus for 6 voltages, 6 pairs of distribution curves for the *intercepted* scattered and primary beams were obtained, the ratio of the areas of each pair giving the value of S'/P' for each voltage. Al and An were used as absorbers. In both cases, the computed curves (not given) for S'/P' did not agree with the exp't. results of B and K. E. J. C.

A note on the Compton effect. SATYENDRA RAY. *Proc. 15th Indian Sci. Cong.* 1928, 69 E. J. C.

Röntgen-ray equipment of the Institute for Experimental Physics of Royal Joseph University for Technical Sciences at Budapest. B. POGÁNY. *Technika* 11, 74-80 (1930). S. S. DE FINELY

Quantitative x-ray analysis. Copper-silver and copper-zinc alloys. HENRY TERRY and ERIC GEORGE VICTORY BARKETT. *J. Phys. Chem.* 35, 1156-67 (1931) —

The intensities of the $K\alpha$ radiation in the x-ray emission spectra of Cu-Ag alloys are detd. by measurement of their ionizing power, and of Cu-Zn alloys by both the ionization and the "wedge camera" methods. The compn. of an alloy can be detd. quantitatively by these intensity measurements only when (1) the elements in the alloy are adjacent to one another in the periodic system, and (2) the emission spectrum of any one constituent metal does not fall within the absorption region of any of the others. The $K\alpha_1$ line of Ag falls within the absorption region of Cu, thus causing a diminution in the intensity of the Ag radiation and a strengthening in the Cu radiation intensity. No absorption of the $K\alpha_2$ Zn line occurs in the Cu-Zn alloys. For Cu-Ag alloys the relation between intensity and at. concn. may be represented over a wide range of concn. by: $I_{Ag}/I_{Cu} = M(C_{Ag}/(C_{Ag} + C_{Cu}))$ for Cu-Zn alloys the relationship is linear. The so-called "wedge camera" method is developed by the authors for detn. of the relative intensities of lines in an x-ray spectrum, being dependent on the measurement of their penetration of an Al wedge. J. BALOZIAN

The photographs of powders in relation to the physical and chemical purity of the substance. GEORGINA R. LEVI. *Atti III congresso nas. chim. pura applicata* 1930, 118-21.—L. claims that soft x-rays (2.3 to 1.5 Å) are sensitive to surface alterations. Thus Ag₂O incipiently reduced but showing no free Ag by chem. analysis shows Ag lines in the powder diagram. When 70% reduced only the Ag lines are observed. Hard rays (0.75 to 0.60 Å) are less sensitive. For mixts. the powder diagrams give qual. but not quant. results. J. B. ARSTIV

A note on the relative absorption of the primary and scattered x-rays by silver and tin. S. R. KNASTORF. *Proc. 15th Indian Sci. Cong.* 1928, 83.—The results (not given) of a few absorption expts. not only show that the theory of wave length change on scattering is untenable but that the greater absorptibility of the primary radiation, as observed by Gray (*C. A.* 17, 2229) is most certainly due to the f transformation in Ag. E. J. C.

X-ray interference by di- and tri-atomic molecules of light gases. IL. GAJEWSKI. *Physik. Z.* 32, 219-21 (1931).—Scattering expts. were made, with a Cu radiator, on N₂, O₂, CO₂, NH₃, and the vapors of CS₂ and H₂O. Interatomic distances calcd. from the data were compared with those from other types of detns. The exptl. scattering curves were compared with the theoretical curves from the Debye scattering formula. C. F. P. JEFFREYS

X-ray dispersion in copper sulfate crystals. L. JÓZSEF NAGY. *Matematik. és Természettud. Közlem.* 47, 70-7 (German abstract 78) (1930).—The absorption limits of the Cu K line were studied by means of a Seemann app. on 3 different oriented surfaces of CuSO₄ crystals. Not only the grating space but also the dispersion changes with the reflecting surface. Larger dispersion belongs to larger grating space. No dispersion data agreed with the theoretical value of $\lambda = 3 \times 10^{-8}$. S. S. DE FRIALLY

Absorption of x-rays by colloidal solutions. KAILASH NATH MATUR and HARI RAM SARTY. *Proc. 15th Indian Sci. Cong.* 1928, 79. E. J. C.

"Ghost lines" caused by an adsorbed layer on the grating surface. SATYENDRA RAY. *Proc. 15th Indian Sci. Cong.* 1928, 90-1.—The "ghost lines" are supposed to be caused by a periodic irregularity in the grating ruling. To test another theory of the lines R. washed the surface of the grating, while in its place on the spectrometer table for viewing the "ghost lines," by means of alc. and gasoline, resp. In both cases lines more widely spaced than the ghost lines appeared with the film of the liquid on the grating surface, the distance between the lines closing up with evapn. of the film and finally collapsing into the ghost lines as a "limiting position" of sorts. The "ghost lines" for the green line of the Hg lamp were also found with subjective observation to appear violet in color on looking obliquely through the eye piece. E. J. C.

The temperature radiation of thallium vapor. K. MAJUMDAR and P. K. KICHILU. *Proc. 15th Indian Sci. Cong.* 1928, 84.—The emission spectrum of hot Tl gas contained in a vacuum furnace with temp. varying from 1500° to 2300° was investigated with the aid of a small quartz spectrograph. Niemi's ion line was obtained at 1500°. Emission begins at 1800°. At 1900° the $2p_2-3s$ line (5370) was obtained in emission, while $2p_1-3s$, (3775) was obtained in absorption. Even at 2300°, the latter could not be obtained in emission. E. J. C.

Spectra of the helium glow discharge. H. McN. COWAN, W. L. BROWN, K. G. EMELEY. *Nature* 127, 593 (1931).—In the cold cathode glow discharge, the line 4713 of the He spectrum is much feeblar than the 5016 line in the neg. glow, while the 2 are of comparable intensity in the anode glow. From the current voltage characteristics of the probe wire, the distribution of velocities among the electrons is found to be not Maxwellian in the neg. glow. It can be fairly well represented by the superposition

and 7 satellites, resp. The line 6907 consists of not less than 5 components and 6231 of not less than 7. On account of the complicated structure the measurements made on the patterns obtained in the 2 quartz 1 mmmer plates have not been enough to fix the wavelength intervals of these 2 lines without ambiguity. By assuming that the fine structure of Hg lines is due to quantized nuclear rotation and taking the resonance line 2510 Å 11 to consist of 5 components, the ^{199}Hg term gives only one fine structure level and the $2P_{1/2}$ level splits into 4 levels $n-1$, n , n and $n-1$, where n is the nuclear spin-moment in units of $\hbar/2\pi$. Hence there can only be 4 components for 2530 Å U on the assumption that all atoms of Hg have the same nuclear moment. G. C.

Fine structure in the mercury singlet terms. S. L. ANSKY. *Proc. Roy. Soc. (London)* A130, 558-78(1931). The structures of singlet and intercombination lines in Hg were investigated to obtain information as to the structure of singlet terms. A high frequency electrodeless discharge in Hg vapor at a pressure of 0.001 mm. was used because in this source the singlet combinations are produced with high intensity without reversal. The observations were made with a Fabry Perot interferometer, with plate separations of 1 to 100 mm. crossed with a Hilger P_{10} spectrograph. The following structures were found: $1S$, $8P_{1/2}$, sextet; $1S$, $8P_{3/2}$, septet; $1S$, $8P_{1/2}$, octet; $7S$, $8P_{1/2}$, septet; $6P_{1/2}$, $8S_{1/2}$, quintet; $6P_{3/2}$, $8S_{3/2}$, triplet and $6P_{3/2}$, $10S_{3/2}$, doublet. Information as to the structures of the $1S$, $7S$, $6P_{1/2}$, $8P_{1/2}$, $8P_{3/2}$ levels is derived from the line structures. The structures found for $1S$ $1/2$ lines necessitate more than one value of the spin quantum no., s , of the nucleus and this conclusion is supported by the fact that the $7S$ level is at least double. The septet character of $7S$ $8P_{1/2}$ can be accounted for by s values of 1 one of which is 11 the other 2 being greater than $1/2$. The 0 value is probably associated with the even isotopes (198, 199, 200, 202, 204) and the other 2 with the 2 odd isotopes (199, 201). W. P. MUMFORD.

Lommel's analysis and continuous spectra. NATHAN N. ROSE. *Proc. 15th Indian Sci. Cong.* 1928, 70. Continuous spectra can be explained by resorting to Lommel's analysis, which is capable of giving a continuous train of waves corresponding to a monochromatic radiation. No details are given. P. J. C.

Rotational structure of light molecules. W. W. WATSON. *Physik. Z.* 31, 880-2 (1930). The rotational energy, which may be obtained from the band structure, indicates the electron configuration of light molecules. LOUIS WATKINSON.

The anomalous dispersion of crystalline lead chloride in the region of its first ultra-violet frequency. KARL HILTZ. *Z. Physik* 66, 339-11(1930). H. B. J.

The resonance spectra of sulfur vapor. J. VANIERENHOEF. *Comp. rend.* 192, 737-9(1911).—B vapor excited by the H γ arc lines at 2988, 3129 and 3142 Å 11 shows a resonance spectrum consisting of shoulders or spritzes in the blue and green. W. P. MUMFORD.

Molecular spectra of mercury, zinc, cadmium, magnesium and thallium. H. HAMADA. *Nature* 127, 555(1911). W. P. MUMFORD.

Magnetic susceptibility and absorption spectra of complex cyanides. I. HANABASHI P. OXENFORD. *Gazz. chim. Ital.* 60, 916-57(1930), cf. Cambi and U., *C. A.* 21, 3021, 3123. In all Fe complexes studied up to now there is a close relation between the magnetic and the optical properties, i. e., in the paramagnetic state the absorption spectrum is richer in bands than is that of the diamagnetic complex. Other cyanides, including those of Cr, Mn, Cu and Ni, have now been studied. Measurements of the magnetic susceptibility were made by the method of Gouy as modified by Schaffer and Taylor (cf. *C. A.* 20, 1762), and of the absorption spectra with the apparatus described (cf. *C. A.* 25, 1524). $\text{K}_2\text{Cr}(\text{CN})_6$ is paramagnetic (cf. Hiltz, *C. A.* 22, 4017), with 10 Weiss magnetons, and it has 2 absorption maxima at $\nu \times 10^{-11} \approx 800$ and 1150, resp. (the latter of great intensity). $\text{K}_2\text{Cr}(\text{CN})_6$ was too unstable to obtain such reliable data, but its magnetic moment is greater than that of $\text{K}_2\text{Cr}(\text{CN})_6$. $\text{K}_2\text{Mn}(\text{CN})_6$ and $\text{K}_2\text{Mn}(\text{CN})_6$ are paramagnetic with 16 and 10 magnetons, resp. (cf. Hiltz, *loc. cit.*; Freed, *C. A.* 24, 4700). The former shows a strong band at $\nu \times 10^{-11} \approx 924$, and $\text{K}_2\text{Mn}(\text{CN})_6$ a band at $\nu \times 10^{-11} \approx 1271$. The extreme instability of $\text{K}_2\text{Mn}(\text{CN})_6$ (cf. *C. A.* 22, 2522) made it impossible to measure its magnetic and optical properties. $\text{K}_2\text{Co}(\text{CN})_6$ has electronic properties like those of $\text{K}_2\text{Cr}(\text{CN})_6$; it is diamagnetic, and its absorption spectrum has 2 maxima, $\nu \times 10^{-11} \approx 975$ and 1152, with low ϵ values. Contrary to Hiltz, $\text{K}_2\text{Co}(\text{CN})_6$ is diamagnetic, in spite of the odd no. of electrons in the central atom, which gives origin to a magnetic spin (cf. Lewis, *C. A.* 18, 3513). $\text{K}_2\text{Co}(\text{CN})_6$ is more active optically than $\text{K}_2\text{Cr}(\text{CN})_6$ and has 2 absorption maxima at $\nu \times 10^{-11} \approx 932$ and 1070, with ϵ values 10 times greater than those of $\text{K}_2\text{Cr}(\text{CN})_6$. Accordingly, a notable variation in the optical behavior does not correspond to any appreciable difference in magnetic moment. An electronic explanation of this incon-

gruity is offered. The behavior of Ni complexes is still more complex than that of the Co compounds. $K_2Ni(CN)_4$ is an exception to the general rule (cf Stoner, *Magnetism and Atomic Structure*, C. A. 21, 859), since it is isoelectronic with $K_2Mn(CN)_4$. Neither does the magnetic susceptibility of $K_2Ni(CN)_4$ vary, and it is diamagnetic in spite of 33 electrons, which by theory should mean a strongly paramagnetic compound. To explain these phenomena, different hypotheses are considered. Other studies of various Ni complexes make it doubtful whether these incongruities depend only upon the central atom. That the diamagnetism does not result from the same causes as that of $K_2Fe(CN)_6$ and $K_2Co(CN)_6$ is proved by the absorption spectra. $K_2Ni(CN)_4$ shows 3 bands $\nu \times 10^{-12} = 962, 1050$ and 1125 , with very high ϵ values. The same maxima appear in the spectrum of $K_2Co(CN)_6$, which shows a notable variation in the absorption only in the visible field, where the extinction is approx. 10 times that for $K_2Ni(CN)_4$. In general, the magnetic properties of neither the cyanides of Co^{++} nor of Ni^{++} and Ni^{+} satisfy the electronic theory which in the past has been applied to these complexes, and there are great differences between the magnetic and optical behavior in that a diamagnetic state corresponds to great optical activity. Other complexes are being studied.

C. C. DAVIS

The mercury band system in the neighborhood of the resonance line. RAYLEIGH. *Nature* 127, 125(1931).—In spectrograms of a water-cooled Hg arc the usual bands at 2540 and 2536.52 Å U were observed, additional bands at 2540.37, 2538.44 and 2537.32 Å U shaded from the red were also measured, the bands 2535.82 and 2535.35 on the other side of the resonance line (2536.52) shaded toward the red were also measured.

FRANK MARESH

A new band system of copper hydride. P. C. MAHLANT. *Nature* 127, 557(1931).—A band system consisting of 6 band heads has been found in the region 2000–2200 Å U. It is different in nature from the 2 systems previously known (C. A. 24, 3951). The bands are double headed, indicating that their emitter consists of an odd no. of electrons; they have therefore been attributed to the CuH^+ mol. The band structure has been analyzed and assigned to a $^1\Pi \rightarrow ^1\Sigma$ transition.

W. F. MOORE

Band spectra of bismuth hydride. A. HEDMER AND E. HULTHÉN. *Nature* 127, 557(1931).—When a quartz discharge vessel fed with 0.5 amp., 1200 v. d. c. is used with Bi vapor distg. at 900° through a narrow end-on tube, a band spectrum near 4722 Å U is excited with great intensity. This band is composed of P and R branches, the system being assigned to a $^1\Sigma \rightarrow ^1\Sigma$ transition in the BiH mol.

W. P. M.

New bands in the secondary spectrum of hydrogen. DATTATRAYA SHRIDHAR JOG. *Phil. Mag.* 11, 761–80(1931).—A review of the principles underlying the deriv. of the theoretical electronic terms of the H mol. in its various states of excitation is followed by a presentation of the bands which, as previously announced (C. A. 24, 3435), have been found to represent the forbidden transition $3p^2\Pi \rightarrow 2p^1\Sigma$. The vibrational and rotational structures of $3p^2\Pi$ are known from the work of Richardson and Das (C. A. 23, 2096), and those of $2p^1\Sigma$ from Richardson and Davidson (C. A. 23, 2883, 3163, 3405). The bands presented in this analysis satisfy the structures of the initial and final terms.

C. C. KRESS

Fine structure in the hydrogen band lines. O. W. RICHARDSON AND W. E. WILLIAMS. *Nature* 127, 481(1931), cf C. A. 24, 1576, 2049.—By examg. the spectrum of H_2 with a large Hilger quartz spectrograph crossed by a reflection echelon R and W. find that the alternate strong lines of 2 typical bands which end on the $2p^2\Pi_{3/2}$ levels are all double with the weaker component on the long wave length side. The observations establish a correlation between the spectra of H_2 and H^+ .

H. W. LEAHY

A study of the widths of the lines in the B band, due to atmospheric oxygen, in the solar spectrum. R. VAN DER RIET WOOLLEY. *Astrophys. J.* 73, 185–93(1931).—The relative widths of 28 λ 's in the B band (6870 to 6920 Å U) in the solar spectrum, due to O_2 in the ear. a's atm., were accurately measured to det. experimentally the relation between line width and the no. of atoms effective in forming the lines. Measurements were made on spectrograms secured at different altitudes of the sun for $\nu = 1/11$ and $1/4$, where ν is the ratio of the intensity of the light at some point inside the line to that of the continuous background just outside the line. The results show that the line widths are approx. proportional to the $1/4$ power of the no. of atoms in the air path traversed by the sun's rays, which is a small but definite deviation from Unsöld's formula in which the widths are proportional to the square root of the no. of atoms.

C. C. KRESS

Raman effect and ionization of electrolytes. LEONARD A. WOODWARD. *Physik. Z.* 32, 212–4(1931).—Raman spectra for solns. of HNO_3 , H_2SO_4 , CH_3COOH , HIO_3 , and HCl were studied. All except HCl solns. gave intense spectra excited by the

4358 Å. U. line of Hg. The Raman line ascribed to NO_3^- was found to increase to a max. of intensity with increasing concn., then decrease. The H_2SO_4 lines also showed such anomalous intensity effects.

C. E. F. JEFFREYS

Polarization of the Raman spectrum of water. C. RAMASWAMY *Nature* 127, 558(1931)—The Raman spectrum of pure water was studied by the original method of Raman. Polarization photographs were taken and the intensities of the bands were estd. from a plate contg. a series of graded exposures of the Hg spectrum. Conclusion. Three different Raman bands excited by the same incident line are differently polarized, and the degree of polarization appears to increase as the intensity of the band increases.

W. F. MEGGERS

The Raman effect in pure water and in several solutions. R. BRUNETTI AND Z. OLLANO *Atti accad. Lincei* 12, 522-9(1930)—Pure water exposed to the Hg arc light gives 3 bands in the visible region green (5050 to 5150 Å. U.), blue (4645 to 4745 Å. U.) and violet (4135 to 4220 Å. U.), excited by the 4358, 4047 and 3650 Å. U. lines of Hg. The differences in frequency of these bands from the exciting lines are 3225, 3469 and 3589 cm^{-1} . In the presence of Cl^- (HCl 50 and 25%, CeCl_3 2.6 and 1.3 M) the 3225 and 3589 cm^{-1} bands diminish in intensity, while 3469 cm^{-1} appears to be intensified. The following NO_3^- solns. were also investigated: HNO_3 (0.5, 3.5, 2.6), NaNO_3 (9.8), NH_4NO_3 (12.5), KNO_3 (3.96, 1.98), $\text{Ba}(\text{NO}_3)_2$ (0.64), $\text{Al}(\text{NO}_3)_3$ (1.86), $\text{La}(\text{NO}_3)_3$ (1.67), $\text{Ce}(\text{NO}_3)_3$ (3.1, 1.5) and $\text{Th}(\text{NO}_3)_4$ (1.65), all expressed as moles per l. Contrary to the effect of Cl^- , the 3589 cm^{-1} band is intensified, while the 3225 cm^{-1} band fades till it disappears entirely in the more concd. solns.

A. W. C.

Raman effect of hydroxyl ions. LEONARD A. WOODWARD *Physik. Z.* 32, 261-2 (1931)—10.4 N NaOH and 12.6 N KOH were investigated with radiation of wave lengths 3654.8 and 3650.2, 4046.8 and 4358.3 Å. U. and showed a common frequency, $\Delta\nu$, of 3615 \pm 25 cm^{-1} . This agrees with $\Delta\nu$ for solid NaOH of 3630 cm^{-1} . A short discussion is given.

WILLIAM E. VAUGHAN

Energy measurements in the visible and ultra-violet. F. BENFORD AND RACHEL F. HOWE *Trans. Illum. Eng. Soc.* 26, 292(1931)—The photographic, the filter and the monochromator methods of measuring energy throughout the visible and ultra-violet spectra are described. The inherent limitations of each method for sunlight and some of the more common sources of light are discussed.

W. T. CLARK

The absorption and scattering of light in opal glasses. G. M. DREOST *Phil. Mag.* 11, 801-46(1931)—Absorption and scattering by a mastic sol with acid fuchsin added was studied since it is not possible to vary κ/σ for glasses, κ = absorption coeff., σ = scattering coeff. Intensity measurements were made with a Glan spectrophotometer by substituting a photographic wedge for the Nicol prism, and thereby decreasing the time and labor involved and increasing the range of intensity ratios determinable. For a parallel incident beam and over a wide range of concn., the equation $I = I_0 e^{-\kappa x}$ was found to apply, where I = transmitted light, I_0 = incident light, x = concn. on an arbitrary scale and σ includes the thickness of the layer. For optically thin layers the coeffs. of scattering and absorption affect the directly transmitted light independently and the equation becomes $I = I_0 e^{-(\kappa + \sigma)x}$. For optically thick layers at high concns. more light is transmitted than given by the exponential formula, due to the reinforcement of the direct pencil of light by repeatedly scattered light. The repeatedly scattered light is detd. by measurement of the light intensity at several small angles to the normal. The proportion of repeatedly scattered light diminishes as the direction of the parallel rays becomes more definite and measurements carefully made in the same direction. Schuster's theory is extended for milk glass with particles of the order of dimensions as the wave length of light. For the thick layers the variation of transmission with the optical thickness was found to be exp'tly. and theoretically the same for a parallel as for a diffuse incident beam. The advantage of the diffuse beam is the greater intensity obtainable. The ratio of κ/σ for Leerdam opal glass hulbs and Philips Argenta hulbs was found to be small.

ARTHUR FLEISCHER

The Becquerel effect of copper oxide in alkaline solution. H. YOGORO KATO AND NAGAO HAYAMI *J. Chem. Soc. Japan* 52, 8-19(1931)—A study was made of the Becquerel effect of Cu_2O and CuO prepd. by the authors' special method, in alk. soln. The time effect of the Becquerel effect and change of wt. of CuO in the soln. when it was and was not exposed to light were detd. The Becquerel effect is said to be due not to photoelectricity but to photochem. change. The decrease of the wt. of CuO when it is exposed to light is proof of this conclusion. The Becquerel effect changes with time and for this reason the results of other investigators do not agree.

R. SAITO

The theory of anisotropic liquids. XIII. The optical activity of twisted structures. C. W. OSBORN *Arkiv Mat. Astron. Fysik* 22A, No. 17, 12 pp (1931), cf. C. A. 24,

3409, 25, 10—Math. The analytical soln. of the problem of the optical properties of twisted structures in the spectral region outside the limits of indescendent colors is presented. Twisted structures possess no rotatory power in the ordinary sense. An incident linear polarized light wave after transition through a layer leaves as an elliptically polarized wave. The usual method of measuring the optical rotation of a nematic substance, by measuring the rotation of a very thin layer and reducing to a 1-mm layer is not applicable to substances with twisted structures. The optical activity of a twisted structure depends on the angle which the vibration direction of the incident light makes with the direction of the mol axis in the lower boundary layers. XIV. The possible structures of nematic substances. *Ibid* No 18, 10 pp.—Math. A. F.

Light emission from alkali halide phosphors. WALTER BENNER. *Z. Physik* 66, 311-27 (1930)—15% to 40% of the light quanta absorbed from the long wave length absorption band appear in the fluorescence band of a TiCl_3KCl phosphor. The max. of this band lies at $\lambda = 300 \text{ m}\mu$. The missing portion is to be sought in a visible emission band with a max. around $470 \text{ m}\mu$. The short wave length absorption band gives similar results. At high TiCl_3 concn. there is a retention of 3.5-5.0% of the absorbed light quanta. No photochem. discoloration was observed. S. BRADFORD STONE.

New laws of phosphorescence. P. L. LADOUCH. *Mathematik its Termizmetul* *Fizika* 47, 79-95 (German abstract 96) (1930)—Photocopy exps. with gelatinous soln. of Rhodamine B proved that its phosphorescence bands are polarized. The phosphorescence band emitted by the dye mol. is partially polarized. S. S. DE LINALY.

The relation of the intensity of fluorescence to the concentration in the case of solid solutions. A. A. DIXON. *J. Optical Soc. Am.* 21, 250-8 (1931)—Measurements were made for liquid and solid solns. of rhodamine B in collodion 4 parts of alc. to one of collodion being used. The app. and exptl. procedure with the Hg arc and W lamp as exciting sources are described. Results show that for both liquid and solid solns. the intensity of fluorescence increases exponentially with a decrease in concn. There is evidence of a const. value for the fluorescence power at low concns. The value for liquid solns. is about 20 times as large than for solid solns. H. W. WALKER.

Fresh evidence in support of the chemical theory of fluorescence. S. S. BHATTAGAR AND KASHINA GOPAL MATHUR. *Proc. 15th Indian Sci. Cong.* 1928, 141-2—In support of Perrin's view of the chem. nature of fluorescence, small changes in the n s of eosin, fluorescein and acriflavine in their aq. solns. were observed on exposing them to a coned beam of light from an arc lamp. The changes in the n s were detd. by means of the Rayleigh interference refractometer, by which a change in n of 1 in 100,000 could be ascertained with accuracy. Expts. were performed at various mol. concns., and the values obtained for the changes in the n s were found to diminish at higher concns. as the intensity of the fluorescent light decreased. In the presence of gelatin and NaOH the values for the changes were again observed to diminish, evidently on account of a retarding influence on the photodecompn. of these fluorescent substances. The values are not given. T. J. C.

Colors of inorganic salts. M. N. SAHA AND S. C. DEN. *Nature* 127, 485 (1931) cl. C. A. 24, 2377—In continuation of previous work S. and D. have measured the absorption spectra of CrCl_3 and FeCl_3 between 1000° and 1400° in a vacuum furnace. CrCl_3 was found to yield a no. of bands at 4100 \AA U. and 4350 \AA U. These bands are ascribed to Cr^{3+} and Fe^{3+} and are due to the magnetic transitions in the d^3 and d^5 shells of the elements. The continuous absorption from 3000 to 2200 \AA U. is ascribed to Cl^- . H. W. WALKER.

The absorption of aqueous solutions of tartaric acid. G. BAUHAUT. *Compt. rend.* 192, 489-90 (1931)—According to B's best results these solns. follow Beer's law. The careful measurements of Lucas and Schwob (C. A. 25, 2052) indicate a weaker absorption and an increase in mol. absorption with decrease in concn. Both sets of results contradict the conclusion reached from the classical hypothesis that 2 forms of tartaric acid exist in soln., namely, that the mol. absorption in the region 2000 - 1500 \AA U. should diminish as the concn. diminishes. JACQUES I. AUSTIN.

The absorption of liquid and solid solutions of rhodamine B in a mixture of alcohol and collodion. A. A. DIXON. *J. Optical Soc. Am.* 21, 259-61 (1931)—By plotting values of k from the equation $I = I_0 e^{-kx}$ (where I_0 and I are the intensities of the incident and emergent light, x is the thickness of the absorbing layer, c the concn. and k the absorption const.) against λ , absorption curves for liquid and solid solns. at different concns. are obtained. λ varies between 5100 and 5900 \AA U. In each case there is a shift in the crest of the bands toward longer wave lengths as the concn. is increased. H. W. WALKER.

Electron diffraction in 1,2-dichloroethane. ERYST BERGMANN AND LEO ENGEL.

Physik Z. 32, 263-4(1931)—Comment on work of R. Wierl (*C. A.* 24, 3439), who claims that 1,2-dichloroethane is a mixt. of 2 stereoisomers, and comparison with evidence to the contrary.

WILLIAM E. VAUGHAN

The absorption of the carbonyl chromophore in the short-wave ultra-violet region. H. LEV AND B. ARENOS. *Z. physik. Chem., Abt. B.* 12, 132-8(1931)—A trim in the absorption curve of pure acetone was found at 2115 Å. In a hexane soln. of acetone, the max. of the short wave band was located at 187 mμ. Water causes displacement toward short waves. The photolysis of acetone became pronounced below 200 mμ. The long wave band cannot be attributed to the C electrons of the chromophore. The increase in photolysis below 200 mμ appears to confirm the ideas of Wolf (*C. A.* 23, 4189). The results also reopen the discussion of results obtained with compounds containing 2 chromophores, i. e. carbonyl and ethylene groups.

FRANK URBAN

Significance of the structure of the hydrocarbon residue for the velocity and equilibrium position in organic reactions. W. HEROLD AND K. L. WOLF. *Z. physik. Chem., Abt. B.* 12, 194-205(1931) cf. *C. A.* 23, 4189; Meerwein *C. A.* 23, 97—The displacement of the absorption band of a chromophore with respect to its normal position or the displacement of the absorption bands of the group itself is taken as a measure of the effect of the structure of hydrocarbon chains on reactions of org. dipolar mols. The dipole moment of the group may be increased or a spatially enclosed or electrostatically screened dipole may be set free to increase activity to form suitable complexes. Various ketones are studied in MeOH and in aq. soln. and various alcohols in EtOH and Me₂CO soln. The constants of esterification and saponification of a no. of esters are also correlated with absorption data, as are also constants of various other reactions. It is shown that a variation of structure of solvent mol. which by formation of solvation complexes, changes the velocity and equilibrium position in org. reactions also exerts a similar influence on the displacement of the absorption bands. Comparison of various series of measurements of this kind with the spectroscopic data in regard to spatial relations of groups should give us a better insight into the space packing of groups.

O. E. S.

The photochemical reaction of iodine with hydrogen. NOBUJI SASAKI AND KENJI NAKAMURA. *Anniversary Vol. Dedicated to Masumichi Chikashige* (Kyoto Imp. Univ.) 1930, 299-302—Preliminary results on the formation of HI by ultra-violet rays of 1800 Å. U. are reported. No trace of I₂ production could be detected with rays longer than 3100 Å. U. It is possible that H atoms are produced by collision of the 2nd kind between H₂ and I₂ excited by radiation of λ about 1800 Å. U. These H atoms may then react with normal I₂ mols.

ALLEN S. SMITH

The chemical action of ultra-violet light on the alkyl iodides. GUY Emschwiller. *Compt. rend.* 192, 799-802(1931), cf. *C. A.* 24, 5634—The action of ultra-violet light on the liquid alkyl iodides which contain 4C atoms or less liberates I and produces a gas. MeI gives CH₄ as the gas, EtI, C₂H₆ and C₂H₄, PrI and iso-PrI, C₃H₈ and C₃H₆, *n*-BuI, *sec*-BuI and iso-BuI, C₄H₁₀ and C₄H₈, *tert*-BuI, C₄H₁₀, C₄H₈, C₄H₆ and C₄H₄. The results are obviously similar except for MeI and *tert*-BuI. The decomposition usually gives I, an olefin and a satd. hydrocarbon, it may be formulated thus: 2RI → (R-H) + (R-H) + I₂. A mol. of RI loses an atom of H and of I, and the R-H rearranges to give olefins. The atoms of H and I act, singly or in combination, on a 2nd mol. of RI to give R and I₂. These photochem. decomps. are interpreted to reveal the existence of a H atom of singular properties in the alkyl iodides, attached to the same C as is the I, except, of course, in the tertiary iodides. The behavior of this H atom is intimately related to that of the I atom. This conception appears to account for certain properties and complex reactions of the alkyl iodides, especially in the syntheses of Wurtz.

R. H. LOMBARD

The photochemical decomposition of nicotine salts. I. PLOTNIKOV AND K. WEBER. *Chem.-Ztg.* 55, 237-9(1931)—Aq. solns. of nicotine, as free base, salicylate, tartrate or malonate, were treated with org. and inorg. substances known to be catalysts for photochem. reactions, and exposed to sunlight. Many substances activate the decomposition, particularly K₂CrO₄, AuCl₃, Fe(NO₃)₃, I and chlorophyll. The same substances are active toward the tobacco itself, although the rate of decomposition is much less. Orange-red light (O. G. filter from Schott & Gen.) was most active in the decomposition of nicotine in tobacco.

GEORGE CALINGAERT

Hydrolysis of acetone in ultra-violet light. M. QURESHI AND N. A. TAHER. *Nature* 127, 522(1931)—For the hydrolysis of acetone in ultra-violet light a direct proportionality has been found between incident light intensity and reaction velocity. When an aq. soln. of acetone in a quartz vessel is illuminated by a Hg arc, HCHO as well as AcOH has been found as a hydrolysis product.

E. J. ROSENBAUM

The photochemical decomposition of nitrous acid. K. S. MURTY AND N. R. DHAR.

J Indian Chem Soc 7, 885-90(1930)—The kinetics, temp coeff and quantum efficiency of the photochem decompn of HNO_3 were studied at wave lengths of 4725, 5050, 7304 and 8500 Å. U., according to the method of Bhattacharya and Dhar (*C A* 23, 2887). Explanations are given for the increase with temp of the temp coeff of the thermal reaction and the fact that the temp coeff of the photochem reaction is less than unity. The velocity of the decompn of HNO_3 is proportional to the cube root of the light intensity. The Einstein equivalence law does not hold for this reaction.

E. J. ROSENBAUM

The function of Br₂ and I₂ ions and the influence of Cl⁻ ions in some oxidation reactions in light. R. M. PURKAYASTHA *J Indian Chem Soc* 7, 691-6(1930)—The kinetics and mechanism of the reaction of lactic and mandelic acids with Br₂ in the presence of KBr were studied. Zero mol constts were obtained when radiation of wave lengths 542 488 448 408 or 365 microns was used. Cl⁻ ions were found to retard the reactions. The semi mol constts obtained by Dhar (*C A* 23, 5414) for the oxidation by I₂ of the Na salts of org acids in the presence of KI are interpreted as resulting from combined light and thermal effects. The light reaction is calcd. to be of zero order.

E. J. ROSENBAUM

Change of susceptibility of paramagnetic salts under the influence of light. D. M. BOSE AND P. K. RAHA *Nature* 127, 520-1(1931), cf *C A* 24, 3453—From theoretical reasons it can be expected that activation of an ion of the transition group elements leads to a lowering of the magnetic susceptibility of the soln. B and R obtained qual. evidence of a definite lowering of the magnetic susceptibility when light of wave length 4338 or 5790 mμ was allowed to fall on a green CrCl_3 soln.

E. J. ROSENBAUM

The action of light on silver bromide. P. S. MACMAHON AND A. C. CHATTERJI *Proc. 15th Indian Sci Cong* 1928, 138-9—Pure AgBr was sealed in a tube containing ppt of Au as absorbent for Ir in an atm of dry O and exposed to sunlight for lengthy periods. Absorption of O in all cases took place in approx the same amt. as that observed in the case of AgCl. It is suggested that the phenomenon is due not to photochem synthesis of an oxy compd, but to the direct absorption of O by the dispersed Ag produced in the photochem decompn of the Ag halide.

P. J. C.

Ultra-violet absorption spectra of the nitriles and amides of 2-methyl-3-pentene. A. CASTILLE AND E. RUFFOL *Bull sci acad roy Belg* [5], 16, 1120-33(1930)—The absorption spectra of the 2 geometrically isomeric 2-methyl-3-pentene-5-nitriles (I) in hexane soln and their corresponding amides (in alc soln) prepd. by Bruylants and Minetti (cf *C A* 25, 2969) were detd and the curves plotted. *trans*-I absorbs more than *cis* I, while the values for their respective acid amides are reversed, that for the acid amide of *cis* I being greater. The crotonic nitriles and acid amides behave in this same fashion (cf *C A* 20, 708).

OPPER F. SHEPPARD

The absorption of radiation in the lower atmosphere and the amount of ozone. CH. FABRY AND H. BUISSON *Compt rend* 192, 457-61(1931)—O₃ present in the atm may absorb the radiation of short wave length from the light of the stars. The value 0.0022 cm of O₃ per km. of air at atm pressure, calcd from optical absorption measurements, is of the same order of magnitude as that obtained by chem analysis. The higher layers of the atm must be richer in O₃, since measurements of absorption give values 20 times the amt calcd on the assumption that all layers contain the amt. given above.

JANET F. AUSTIN

Formation of hydrocyanic acid in the electrical discharge. KURT PETERS AND HANS KOSTER *Brennstoff Chem* 12, 122-7(1931), cf *C A* 25, 1737—Yields of 70% HCN have been obtained when 1:1 mixts of CH₄ and NH₃ were subjected to the elec. discharge at diminished pressures and rapid gas-flow rates. Increases in the reaction yields resulting from increases in discharge potentials agree with the calcd thermal eqvs. With 15% or less of either CH₄ or NH₃, the yields of HCN become quant. when calcd. for the smaller constituent of the mixt. Under high charges, excess CH₄ and NH₃ form C₂H₂ or decompose into N₂ and H₂ resp. With a 3:7 CH₄:NH₃ mixt. a good yield of pure cryst. NH₄CN results. The addn of NH₃ in the C₂H₂ formation from CH₄ prevents hydrocarbon formation without diminishing the yield of C₂H₂, both tar and hydrocarbons reacting with NH₃ to form HCN. The influence of elec. charge, gas flow rate and compn. upon HCN yields is shown in tables and diagrams. The reaction mechanism is discussed at length.

F. W. JUNG

The decomposition of hydrocarbons in the positive-ray tube. H. R. STEWART AND A. R. OLSON *J Am Chem Soc* 53, 1236-44(1931)—By the use of a specially designed tube it is shown that the decompn. of the hydrocarbons in pos.-ray analysis is due to dissociation by the ionizing electrons or to the secondary reaction between ions and neutral mols. Previous expts with H₂ and C₂H₂ gave ambiguous results, since

part of the products reported were found to decompose in the process of analysis. Selective absorption of hydrocarbon ions by propane and butane is observed. H. W. L.

Chemical action in the glow discharge. VI. The oxidation of carbon monoxide. A. KRITH BREWER AND PAUL D. KILICK. *J Phys Chem* 35, 1281-1302 (1931), cf. C. A. 25, 32.—The oxidation of CO in a glow discharge has been carried out at pressures as high as 30 mm of Hg. In the neg glow the rate of oxidation is proportional to the current and nearly independent of the pressure. It is increased by an excess of CO and retarded by an excess of O₂. In the pos column the rate is proportional to the current and independent of the pressure below 10 mm, but at higher pressures increases to the ignition point. CO₂ synthesized in the neg glow reaches the wall as pos ions, that formed in the pos column is deposited as neutral moles. The bands of the first neg group of CO appear with about equal intensity over a wide pressure range. The data are interpreted as indicating a simple ion cluster mechanism for the neg glow and an ion chain mechanism for the pos column. CO⁺ ions appear much more effective in initiating the reaction than O₂⁺ ions. An estimate of the rate of production of pos ions in the neg glow yields a ratio of M/A of 2 for the no. of moles synthesized to pos ions termed VII. The dissociation and oxidation of methane, *Ibid* 1293-1302.—CH₄ is converted quantitatively into C₂H₂ and H₂ in the glow discharge at liquid air temps. Approx. one C₂H₂ mol is formed per CH₄⁺ ion. CH₄ in a 1:2 CH₄:O₂ mixt is oxidized completely to CO₂ and H₂O. Reaction seems to be initiated primarily by CH₄⁺ ions, being accelerated by excess CH₄ and retarded by excess O₂. The max rate of CH₄ oxidation occurs with a 1:1 CH₄:O₂ mixt. The oxidation reaction is most pronounced in the neg glow, negligible in the dark spaces and small but fairly const throughout the pos column. The efficiency of the pos column in terms of electron volts per mol increases rapidly with the pressure above 5 mm, indicating the presence of ionic chain reactions. P. D. KILICK.

I electrolytic phenomena in oxide-coated filaments (BRECKER) 4. Absorption spectra of the fractions of serum albumins (PAITIN) 11A. Magnetism and molecular structure (BHATTAGAR, BHATIA) 2. A low-density Cd-vapor lamp (VENKATASACHAR) 4. Reactions at the surface of hot metallic filaments (SRIKANTAN) 2. Irradiating medicaments, foodstuffs, etc (Austrian pat 151,657) 17.

BLUMBAUM, IRMA. Quantitative Strahlungsmessungen an künstlichen und natürlichen Strahlungsquellen. Jena Fischer 88 pp M 6.

HAAS, ARTHUR. *Quanta et chimie*. Translated from the German by Jeanne Perrenot and P. Leflangon. Paris Gauthier-Villars et Cie 60 pp F. 15.

Electron-emission element. SAMUEL RUBEN (to Ruben Patents Co). U. S. 1,799,645, April 7. In a low-pressure elec device which functions through an elec. discharge, a cathode is used contg a mixt of ferrosoferrie oxide, chromic oxide and SrO, fused and cast into suitable form. Cl. C. A. 24, 757.

Identifying crystals such as precious gems by optical analysis of emitted rays. FAUSTE F. READ. U. S. 1,799,601, April 7. App. is described.

4—ELECTROCHEMISTRY

COLIN G. FRICK.

A survey of electrochemical industries of Japan. YOGOCO KATO. *Trans. 2nd World Power Conference* (Berlin) 1, 306-10 (1930).—Brine is electrolyzed to form NaOH, bleaching powder and bleaching liquor. By product H₂ is used for the hydrogenation of oil from fish and soy beans. H₂ from the electrolysis of H₂O is used for NH₃ synthesis, O₂ for synthetic AcOH and for oxyacetylene welding. Cu and Fe are purified electrolytically. Electrolytic Zn is produced. Metallic Na is made by the electrolysis of fused NaOH. Au, Ag, Ni and Cr are electroplated. The electrothermic industries are occupied in the production of CaC₂, cyanamide, ferro Si, ferro-Mn, ferro Cr, elec. steel, etc. Synthetic AcOH is made from acetylene. Methanol production is partly commercial. Pb storage batteries and dry cells are made. ALDIN H. FERRY.

The new electrochemical laboratories of the University of Grenoble. R. SEVIN. *J. four Elec* 40, 45-8 (1931), 2 illus. C. G. F.

The new carbide-cyanamide plant at Wingles (France). RAYMOND SEVIN. *J.*

four elec 40, 51-8(1931), 10 illus.—The two 6000 kw carbide furnaces employ the Mignet electrode. Details of installation are given. C G F.

Industrial electric heating XV. The heating and cooling of metals. N R STANSEL AND S L HOYT. *Gen Elec Rev* 34, 302-9(1931), cf C A 24, 4225—A review. C G F.

Two-thousand-kilowatt seventy-five-ton furnace anneals large castings. A G HOTCHKISS. *Flec World* 38, 768(1931)—A 2000 kw car bottom type elec annealing furnace with a max loading capacity of 75 tons has heating units of Ni Cr ribbon resistor type. W. H. ROYNTON.

Automatic electric apparatus for the sterilization of solutions by direct heat without alteration of the titer. H ANGELT. *Boll chim farm* 69, 804-10(1930). G S.

The determination of the transport numbers of metals on electrolyzing their alloys. RICHARD SIMONCE AND RICHARD IRINA. *Monatsh* 57, 112-64(1931)—Molten alloys were electrolyzed by the method of Kremann, Muller and Kienzl (C A 19, 5, 1979). Longer tubes were used in order to obtain an intermediate layer of unchanged compon. The 19 alloy pairs investigated could be divided into 3 groups. With K Hg, Pb Hg, Sn Hg, Sn Pb, Sn Sb, Bi Sn, Bi Cd and K Na, one of the metals migrated toward the anode and the other toward the cathode. With Na Hg, Bi Hg, Sn Zn, Sn Al, Ag Al, Ag Bi, Zn Sb and Sn Cd, either this same behavior was noted, or else both metals migrated toward the anode, depending upon exptl conditions. No definite relation could be established between exptl conditions and the direction of migration except that frequently migration of both metals toward the anode took place at higher c ds. With Ag Sn, Ag Sb and Pb Na, both metals migrated toward the anode. It is concluded that in order to det transport nos for molten alloys such as Hittorf obtained for aq solns, many sources of exptl error must be eliminated. C L W.

The galvanic tension of lead-gold alloys. FRANZ GRINGL AND ROBERT BAUM. *Monatsh* 57, 165-76(1931)—The galvanic tensions of the cells $Hg | HgCl | satd KCl | 0.1 N HCl | satd with PbCl_2 | Au, Pb_{(x)}$ were measured at 2°C on an Ostwald capillary electrometer. With increasing Au concn, the potential remained practically that of Pb at 40% Au it dropped and attained a min at 60% Au, then rose gradually to that of Au. It is concluded that the potentials of the 2 compds $AuPb_3$ and Au_2Pb lie close to that of Pb. CURTIS L. WILSON.

The cathodic behavior of pyrite and chalcopyrite. ATSUSHI MATSUBARA AND JITSUTARO TAKUBO. *Anniversary Vol Dedicated to Masumio Chikashige* (Kyoto Imp Univ) 1930, 311-39(in German)—The elec reactions of Fe and Cu pyrites as cathodes were quantitatively investigated. The app used permitted measurement of the gases evolved and electrolyte used and sampling of the cathode soln. The cathode was prep'd by sealing a crystal of the pyrites in a glass tube. Expts were made at 6 v and at 20 v. Fe pyrites is dissolved in dil H_2SO_4 with the evolution of H_2S by passage of current. If the sample is anodically polarized either dry or in the electrolyte, a relatively large amt of H_2 is evolved in comparison to the quantity of current consumed in the subsequent cathodic polarization in dil H_2SO_4 . Pretreatment by cathodic polarization shows the opposite effect. Cu pyrites is decomposed, forming H_2S , Cu_2S and $FeSO_4$. With preliminary anodic polarization and subsequent cathodic, a relatively small amt of H_2 is evolved and a large amt of the decompn product. The effect on the mineral is opposite with preliminary cathodic polarization. A greater amt of decompn product was found than would correspond to the current consumption owing to the action of the acid. If the cathodic polarization proceeds further, metallic Cu is sepd. A theory of ionization of the 2 minerals is advanced to explain the exptl results. A S S.

The electrolytic preparation of white lead. J F SACHER. *Chem.-Ztg* 55, 189-90(1931)—Various patented processes are reviewed. While the electrolytic process possesses the advantages of avoiding dusting and of being continuous and labor saving, there are serious disadvantages such as the difficulty of producing a purely basic Pb carbonate, higher installation cost, complicated chem control and the danger of discoloration by impurities from the electrodes. It is claimed that the electrolytic process has not attained the reliability of the corroding process, and probably never will (cf the Sperry process (U S 1453,620), C A 17, 2197). CURTIS L. WILSON.

Electrochemistry of magnesium. SVEN BODFORSS. *Z physik Chem*, Abt A, 153, 83-106(1931), cf C A 21, 524, 22, 719—The change of potential of a Mg electrode with time was det'd in $MgSO_4$ and $MgCl_2$ solns. The effect of varying the acid content and varying the buffer in the presence of neutral salts and NH_4 salts or of hydroquinone, resorcinol succinimide, pyrocatechol, CO , CO_2 and H_2 was det'd. Mg becomes less noble as the H^+ ion concn of the soln is increased, till a max is reached.

This is explained by the reaction $Mg \cdot + Mg(ion) + e \text{ or } 2Mg \rightleftharpoons Mg_2(ion) + 2e$

ARTHUR LEISCHER

Direct electrochemical production of permanganic acid from manganous salt solutions. LÁSZLO DE PUTNOKY *Matematik és Természettud Értesítő* 47, 732-48 (in German 749-50)(1930) MnI_2 is formed on the anode during the electrolysis of $MnSO_4$ soln. with a diaphragm in the presence of H_2SO_4 and HCl . The primary product of anodic oxidation is $HMnO_2$, which reacts with $MnSO_4$ and HCl to form MnI_2 . Only after all the Mn^{++} salt has been oxidized to Mn^{+} can the $HMnO_2$ remain as the final product. $HMnO_2$ is then formed by the oxidation of MnI_2 , which therefore decreases in amt. during electrolysis. The MnI_2 could not be detected, probably because it is instantly oxidized. S. S. DE LINÁLY

Series refining of copper and the economics of the process. A. GIZUNOV AND N. CUVARINOV *Chem. Zhurn.* 6, 117-62 6(64) (English)(1931) The multiple and series systems are described and on the basis of the charts given it is concluded that the series system is more economical only when the price of Cu is high and the anodes contain but a small percentage of impurities. J. KILERA

The Nernst equation. I. A new equation for electrode potentials. MARTIN MEYER *Chem. News* 142, 170-83(1931) Some $e.m.f.$ data were graphically examined. A new equation for the $e.m.f.$ of a single electrode or concn. cell is derived on the basis of van der Waals' gas equation. It is predicted that the function $f = [a + b - \phi RT \log(P/p)] / \phi(P-p)$ (ϕ = numerical coeff., R = universal gas const., T = Abs. temp., P = soln. tension of ions, p = soln. pressure of metal) should be a const. for any substance where internal attraction can be assumed to be equal to 0. R. H. CHERRY

Irreversible phenomena of electrolysis (overvoltage and passivity of the electrodes) ERNESTO DI NINA *Atti III congresso naz. chim. pura applicata* 1930, 414-9—Review of electrolysis and polarization phenomena. G. M. MURFIN

Electrodeposition of metals from nonaqueous solvents. I. I. AUDRIETH AND H. W. NIELSEN *Chem. Reviews* 8, 335-52(1931) Electrodeposition of metals from CaH_2 , $HCONH_2$, $AcNH_2$, liquid NH_3 , and other aprotic solvents offers many interesting possibilities. LOUISE KALLIA

Chemicals—their functions in electroplating. L. C. PAN *Brass World* 27, 70-81(1931) The metal compd., the conductor, the anodic depolarizer, the brightening agent and the buffer and the parts they play in electroplating solns., are briefly discussed. W. H. BOYNTON

Recovery of silver and gold from plating solutions. K. BILMAYER *Metallwaren Ind. u. Gahano Tech.* 29, 141-5, 169-71(1931) Waste solns. are disposed of in 3 ways: (1) the solns. are given directly to a ppn. plant, (2) the preliminary ppn. is done at the plating plant and the final recovery at the ppn. plant, (3) the entire work of recovery is done at the plating plant itself. Chem. methods of recovering Au and Ag from acid and alk. solns., and from liquids which contain the noble metals in the form of dust, are described. CURTIS L. WILSON

Work supervision and testing of plating baths. H. KRAUSE *Metallwaren-Ind. u. Gahano Tech.* 29, 145-7(1931), cf. *C. A.* 24, 3145. CURTIS L. WILSON

Electromotive behavior of cupric oxide. BASRUR S. RAO AND N. G. CHOKANNA *Proc. 15th Indian Sci. Cong.* 1928, 115—The half cell $Pt|CuO, Cu_2O, N NaOH$ has been observed to give an inconsistent potential difference by Allmand (*C. A.* 4, 1917) and by Maitland (*C. A.* 20, 1923) no satisfactory explanation has hitherto been available. R. and C. carried out measurements with cells in which the CuO used had been first warmed to dryness with an a/c soln. of $NaOH$ and then heated at different temps. with suitable precautions. Such heating was found to cause an aging effect which tends to give const. potential values. No appreciable aging effect was noticed with Cu_2O under similar treatment. F. J. C.

A low-density cadmium-vapor lamp. B. VENKATESACHAR *Proc. 15th Indian Sci. Cong.* 1928, 67—A Cd lamp with a hot solid Cd cathode and a water-cooled Cd anode was designed. A study of the photographs confirms the results obtained with Hg. Particularly striking is the appearance of the higher members of series lines and of the so-called "forbidden lines" at a pressure of about 2 mm. of Hg. Description of the lamp and detailed results are not given. F. J. C.

Electrolytic phenomena in oxide-coated filaments. J. A. BUCKER *Trans. Electrochem. Soc.* 59 (preprint), 15 pp(1931) cf. *C. A.* 24, 4214—A critical survey of the literature shows that the current through the oxides in oxide-coated filaments is carried by electrons, negative oxygen ions and positive barium ions. The proportion of current carried by each depends upon the exact composition and method of preparation of the oxide coating, on the heat treatment and on previous electrolytic effects. Pre-

sumably the conductivity is greatly affected by Ba and O dispersed through the oxide. New experimental results show that for a particular $\text{BaO} + \text{SrO}$ filament, the cond. was given by $171 \times 10^4 - 173 \times 10^4/T + 5.55 \times 10^{-4} - 0.62 \times 10^{-4}/T$. The current is proportional to the voltage only so long as the current is small, otherwise the products of electrolysis alter the cond. Polarization currents are caused by the Ba and O which are produced by electrolysis. These currents decrease rapidly even at temps. near 600°K , thus showing that Ba and O diffuse at low temps. C G F.

The arrangement of the microcrystals in white Sn deposited by electrolysis (HIRATA, *et al*) 2. The steam and electric power plant of Imperial Chemical Industries, Ltd., at Billingham (HUMPHREY, *et al*) 13. Ni cast iron and its uses in the electrical industry (HIRSCH) 9. Comparison of rust protection of Fe by Zn, by Cd and by Zn-Cu alloys and the electrodeposition of such alloys (MATTHEWS, HARDY) 9. Cu extraction from ore (U. S. pat. 1,799,255) 9. Electric furnace for glass manufacture (U. S. pat. 1,799,371) 19. Sealing composition [for electric condensers] (Can. pat. 309,681) 18. Potting compositions (Can. 311,014) 18. Drying storage battery plates with heated gases (U. S. pat. 1,799,249) 13. Working up cyanide solutions to give K ferrocyanide and the preparation of K ferrocyanide by treatment with compressed air and subsequent electrolytic oxidation (GLOUB, *et al*) 18.

Accumulators. COMPAGNIE GÉNÉRALE D'ÉLECTRICITÉ Fr. 697,390, May 24, 1930. The edges of fibrous separators for accumulators are coated with a substance resistant to the electrolyte.

Accumulators. ERNST BERGMANN. Austrian 121,757, Dec. 15, 1929. Sulfation is prevented by adding $(\text{NH}_4)_2\text{SO}_4$, $\text{Al}(\text{SO}_4)_3$, and a little org. acid to the H_2SO_4 filling. The salts must not be added in alum forming proportions. The org. acid may be added in the form of a salt. A suitable filling contains distd. water 755 and H_2SO_4 240 cc., to which are added anhyd. $\text{Al}(\text{SO}_4)_3$ 18, $(\text{NH}_4)_2\text{SO}_4$ 4 and basic $\text{Al}(\text{OAc})_3$ 5 g. The invention is particularly applicable to starting batteries for automobiles, etc.

Electrodes for accumulators. P. PALÁNYI. Hung. 102,071, March 10, 1930. Positive electrodes for alk. accumulators are made of MgO , $\text{Mg}(\text{OH})_2$, or MgCl_2 mixed with CuO and 3-15% powd. graphite, charcoal, coke or onitracite. Negative electrodes are formed of the above Mg compds. and ZnO contg. a little H_2O . These mixts. are added to powd. resin or pumice, formed into electrodes and heated to redness.

Making electrodes for selenium cells from colloidal graphite. TELEFUNKEN GES. FÜR DRAHTLOSE TELEGRAPHIE M. B. H. (Fritz Michelssoen, inventor). Ger. 522,291, Feb. 14, 1929.

Selenium cell. EMERICH SPIELMANN. Austrian 121,671, Oct. 15, 1930. Structural features are described.

Selenium cells. TELEFUNKEN GES. FÜR DRAHTLOSE TELEGRAPHIE M. B. H. (Otto von Bronk, Fritz Michelssoen and Fritz Schroeter, inventors). Ger. 522,002, Dec. 2, 1928. Se cells sensitive to ultra red rays are formed in an atm. of inert gas. c f A.

Modified Bunsen cell. L. JESZENSKY. Hung. 102,192, Dec. 24, 1929. NaOH or KOH is used as electrolyte with (1) KMnO_4 dissolved in H_2SO_4 contg. water (potential 2.7 v), or (2) Na or K dichromate (2.5 v), or (3) concd. HNO_3 (2.4 v), or (4) H_2SO_4 contg. HgO (1.9 v) or (5) pyrolusite powder moistened with acid water (1.8 v), as depolarizing liquid.

Electrochemical condenser. LÉON SEGAL and HENRI ANDRÉ. Fr. 698,698, Oct. 9, 1929. A Na-alkali (potash or soda) Sn combination is used.

Electroplating apparatus. JOHN J. WIALIN (to General Elec. Co.). U. S. 1,798,094, March 31. Structural features.

Apparatus for electroplating jewelry chains, etc. MAX FESSLER. Fr. 697,885, June 25, 1930.

Electrolytic metal deposition. ERNST KELSEY (Franz Halla, inventor). Austrian 121,986, Nov. 15, 1930. The electrolytic manuf. of metals is conducted under the influence of high frequency sound waves. Cf. C. A. 22, 721, 23, 1573.

Covering metals with aluminum. SOC. ANON. POUR L'IND. DE L'ALUMINIUM. Fr. 698,748, July 8, 1930. Metals such as Fe or light alloys are given an adhering coating of Al, by electrolysis. The base metal serves as cathode, and the electrolyte has a m. p. lower than that of Al. One or more metals more noble than Al is pptd. in small quantities at the same time as the Al at the anode, e. g., small quantities of Pb. Before commencing the pptn. the cathode is allowed to function as anode.

Chromium plating. AUGUSTE HOLLARD (to Soc. nouvelle de l'orfèvrerie d'Ercuis). U. S. 1,799,851, April 7. See Fr. 661,182 (C. A. 24, 28).

Removing mistiness in electrolytic baths during chromium plating. WILHELM STEINHORST. Ger. 504,815, May 25, 1927. The mist is dispelled by a high tension electrostatic field applied by pointed conductors outside the bath.

Chromium-stripping bath. JOSEF F. MCCULLOUGH (to The Ternstedt Manufacturing Co.) Can. 411,131, May 5, 1931. An electrolytic bath for stripping Cr-plated articles, with the article acting as an anode, comprises an aq. soln. contg. 60-120 g. oxalic acid per l. A temp. of 77-85°F. and a potential of approx. 12 v. are used.

Electrodeposition of the metals of the platinum group. BAKER & CO., INC. Fr. 698,153, June 27, 1930. See U. S. 1,779,436 (C. A. 25, 38).

Electrolytic production of metallic powders. EUGENE DROUILLY (to Soc. anon. Trefileries & Laminaires du Havre, anciens établissements Lazare Weiller, Soc. coopérative de Rugles et la canalisation électrique réunies). U. S. 1,799,157, April 7. Metallic powders such as Cu are produced by electrolyzing a soln. of a metal salt such as CuSO_4 contg. the carbonaceous reaction product of glucose and H_2SO_4 substantially in colloidal dispersion.

Electrolytic production of unstable products. ÖSTERREICHISCHE CHEM. WERKE G. M. B. H. Austrian 121,750, Oct. 15, 1930. The electrolytic manuf. of persulfuric acid or its salts is effected in a cell having a diaphragm arranged close to the anode. The anolyte is caused to flow rapidly through the cell in a thin layer, while the main body of electrolyte is cooled and passes slowly through the cell, preferably in counter current to the anolyte. Other electrolytic processes yielding unstable end products are similarly conducted, the diaphragm being located close to the electrode yielding the unstable product. Details of the app. are described.

Electrolytic cell suitable for caustic-soda production from brine. HERBERT I. ALLEN and DOUGLAS B. AYEAST (to Electron Chemical Co.). U. S. 1,798,575, March 31. Structural features.

Purifying solutions. I. G. FARRENHUT A-G. Fr. 697,900, June 25, 1930. Solns. are purified electrolytically by causing a current to act on foreign substances dissolved in low concn. in an electrolyte, the electrolyte being caused to pass through electrodes of metallic gauze, the openings of which are very small. Examples are given of the removal of As from H_3PO_4 , H_2SO_4 , and AcOH , and Hg from AcOH .

Apparatus for electrolytic decomposition of water or other liquids or solutions under high pressure. JACOB E. NOEGGERATH. U. S. 1,799,116, March 31. Structural and elec. features.

Electrometallurgical products. FRANÇOIS C. S. MIZIER. Fr. 698,492, Apr. 7, 1930. Electrometallurgical products are made by using primary materials powdered and agglomerated by any means. The current is brought to the furnace by electrodes made of the material to be treated. The plasticity of the products before cooling is used for molding the electrodes.

Tantalum. CLARENCE W. BALKE (to Fansteel Products Co.). U. S. 1,799,403, April 7. Ta is obtained by electrolyzing a fused bath contg. tantalic oxide and K fluotantalate.

Ozone. EUGÈNE ROYER. Fr. 697,930, Sept. 25, 1929. Ozone is produced under pressure by a magnetic field created by the dielec. liquid of an annular reservoir in a tight chamber having a central metallic mass and receiving the gaseous mixt. under pressure.

Protective layers on metallic or other electric conductors. JONAS H. MELLQUIST. U. S. reissue 18,030, April 7. Reissue of original pat. No. 1,595,675 (C. A. 20, 3271).

Chemical reactions under electric waves. INSTITUT FÜR PHYSIKALISCHE GRUNDLAGEN DER MEDIZIN. Fr. 697,547, June 17, 1930. In chem. reactions under a luminous elec. wave of continuous current, the interior of the reaction chamber is such that the vol. of the space occupied by the cathodic fall is approx. equal to the vol. of the space between the electrodes. The reaction is carried out under reduced pressure.

Rotary electric furnace. GOTTFRIED TRÜMPER. Ger. 522,004, May 1, 1928.

Electric resistance furnace. GOTTLIEB KELLER (to Akt. Ges. Brown Boveri & Cie.). U. S. 1,798,678, March 31. A system of power-input control is described.

Electrically heated furnace suitable for dental work. PAUL MANTERNACH. U. S. 1,799,695, April 7. Structural features.

Iron-coupled electric induction furnaces. HIRSCH, KUPFER- UND MESSINGWERKE A-G. (Manuel Tama, inventor). Ger. 522,198, Nov. 15, 1927. Elec. features are described.

Electric annealing, etc., furnace. **SIEBERT-SCHUCKERTWERKE A-G** (Hans Gerdien and Walter Juhutz inventors) Ger 522,041, Jan 12, 1928

Electrically heated annealing apparatus for coiled iron bands. **HASCH, KUPFER-UND MESSINGWERKE A-G** (Franz Linnhoff, inventor) Ger 522,109, Nov. 25, 1928

Hardening high-speed steel. **HAARLS VACUUMSCHNITZER A-G** and **WILHELM ROHM** Ger 522,003, July 9 1929 Hardening is effected in an electrically heated tube of cast Cr, to which current may be conveyed by water cooled tubes of Ni or Fe welded on to the Cr tube. The tube may be closed at one end and filled with a protective gas

Heating charges such as tungsten filaments and supports in electric furnaces. **LEON C KELLEY** (to General Elec Co) U S 1,799,102, March 31 The charge in the furnace is heated to the furnace temp and is then kept in the furnace while further heated by heat generated in the charge itself (as by a high frequency coil) independently of the furnace temp. App is described

Heat-exchange apparatus suitable for use as a jet heater in electrolytic copper refining. **JOHN P RATHBON** (to Westinghouse Elec & Mfg Co) U S 1,799,391, April 7 Structural features

Packing crucibles of electric metal melting furnaces. **MAGNUS UNGER** (to General Elec Co) U S 1,799,431 April 7 In packing crucibles to prevent leakage, the crucible is surrounded with a layer of granulated, refractory, heat insulating material which is agitated while being applied to the crucible so as to compact it tightly against the crucible

Gas rich in hydrogen from carbide furnace. **KATUZI HIRI** Japan 90215, Feb 9 1931 Gas contg about 40% of H_2 is made by introducing H_2O vapor into the mixt of CaO and C in the CaC_2 furnace. The process has no bad effect on the quality of CaC_2 produced

Head construction for electric resistance detectors for combustible gases. **ALBERT N LUTKESON** (to Union Carbide Co) U S 1,799,977, March 31

Gas purification. **METALLGESellschaft A-G** (Jakobus R Gies, inventor) Ger 516,763, Jan 3, 1930 An elec gas purifying plant in which a gas canal is arranged between two filter groups is described. Cf C A 25,2061

Cleaning gases electrically. **SIEBERT-SCHUCKERTWERKE A-G** (Carl Buff, inventor) Ger 522,166, Dec 5 1928 The cleaning of the insulators, e g, by a rinsing gas, is automatically regulated in accordance with the amt of dust that has settled on them. Cf C A 25,463

Cleaning furnace gases electrically. **METALLGES A-G** (Hermann Fiesel, inventor) Ger 522,340, Apr 19 1929 The electrically neutral gases are supplied to the electrofilter at a temp above 150° , e g, $200-250^\circ$, and at a relatively low velocity generally less than 1 m per sec. The tension in the electrofilter may increase in the direction of flow of the gases. The method is particularly intended for rotary cement kiln gases

Apparatus for electrical precipitation of suspended particles from gases. **HARRY A WINTERMUTH and CARL W J HEDBERG** (to Research Corp) U S 1,798,964 March 31 Structural features

Apparatus for electrical precipitation of suspended particles from gases. **HARRY A WINTERMUTH and CARL W J HEDBERG** (to Research Corp) U S 1,798,511, March 31 Structural features

Precipitating electrode for electric gas cleaner. **"ILGA" ELEKTRISCHE GASREINIGUNG G m b H** Ger 522,479, June 26, 1929

Electrode support for dust precipitators. **HOWARD D SALISBURY** U S 1,798,761, March 31 Structural features

Arc lamp electrode. **STEPHEN W ORNE** (to The Canadian National Carbon Co., Ltd) Can 310,628 Apr 21, 1931 A cord electrode in which the shell is made of a slowly burning form of C, as petroleum coke, contains small quantities of H_2BO_3 with a suitable binder, and the core consists of 10 parts of rare earth fluorides, 60 parts CaI_2 , 30 parts high grade carbon flour and sufficient tar to serve as a binder. Instead of rare earth fluorides other compds may be used as oxides, chlorides, silicides or carbides. Cf C A 25, 885

Incandescent electric-light bulb. **SAMUEL LYMAN** U S 1,798,745, March 31 The base of the bulb has a reentrant portion provided with a silvered reflecting surface. Various details of construction are described

Electric incandescent lamp filament. **THE E FOULKS** (to General Elec Co) U S 1,800,012, April 7 Structural features

Getter for incandescent electric lamp. **EWALD DIETZ** (to The Canadian Westing-

house Co., Ltd.) Can 310,383, Apr 14, 1931 Residual deleterious gases in an incandescent elec lamp are cleaned up by effecting the thermal decompn therein of copper phosphide A getter compn for evacuated incandescent elec lamps comprises cryolite contg about 0.3-40% phosphorus as copper phosphide

5—PHOTOGRAPHY

E. P. WIGHTMAN

Early history of photography. E. STENGER *Phot Ind* 29, 32-5(1931) —English photographic literature from 1819 to 1870 is covered It is essentially a bibliography treated chronologically C. E. MEULENDYKE

Mechanism of the photographic process H. BAINES *Chemistry and Industry* 50, 256-60(1931) A review (without references) is given of the more recent theories of photographic sensitivity and latent image formation, with a brief discussion of some other phases of the photographic process, such as development and fixation F. P. W.

Photographic emulsion for dry plates MICHIO MIYATA AND A. SASAKI *Rep's Imp Ind Research Inst, Osaka, Japan* 11, No 18 (1931) —The method of prepn of the photographic emulsion together with the effect of the concn of halogen salt and gelatin is given F. I. NAKAMURA

Measurement of sensitivity of emulsions and a comparison of the results of the various methods. L. LOBEL AND M. DUBOIS. *Photographie* 18, 85-6, 108-11(1931) —Twenty one emulsions were compared for sensitivity by the methods of Hurter and Driffield, Scheiner, Jones and Russell, and Labussière, and a practical test is made under standardized conditions for min exposure 1 or for the emulsions tested, which were mainly orthochromatic and compared at a gamma 16, the H & D inertia method gave values which agreed best with the practical exposure test, but generally were of lower value W. WESTWATER

Toning of photographic pictures by sulfur and selenium. J. MILDAUER AND F. MASIN *Chem Beob* 6, 17-20(20 English)(1931) —By controlling all methods of direct as well as indirect S and Se toning on images obtained by the development of AgCl, AgBr and AgBr papers, it was found that the direct toning is most readily accepted by images obtained by colored development, e. g., to brown tones A series of new methods, direct and indirect, is described by using a combined bath of thioarsenate and thioantimonate and a very efficient bath with H_2SeO_4 . Thioantimonates, thiomolybdates and thiotungstates cannot be applied for toning, as such baths give permanent colors to the light parts of images The same is true for the sol arsenates and antimonates of Se Colloidal solns of Se do not attack practically the Ag granulation, thus differing from the S solns Images toned by Se can be bleached by using a bath of $KMnO_4 + KBr$, acidified with a trace of HCl , and then retoned J. KUČERA

Conversion of Scheiner speeds into H. & D. M. BILTZ *Phot Ind* 29, 95(1931) —An exact conversion is not possible because of fundamental differences in the 2 methods of detg speeds In a practical way, B has separately detd the 2 speeds for a large no of com plates and films He plotted the Scheiner speeds as abscissas and H & D speeds as ordinates The wide errors resulting from the use of conversion tables are apparent C. E. MEULENDYKE

New photoelectric recording microphotometer (CARROLL, MOSS) 1. Quantitative investigations on the photographic action of α - and H -particles (BLAU) 3. Energy measurements in the visible and ultra violet (BENFORD, HOWE) 3. The action of light on AgBr (MACMAHON, CHATTERJI) 3. Films (Fr pat 693,727) 23.

Color photography. SPICERS, LTD Fr 693,145, June 27, 1930. Photographic prints are obtained in natural colors by exposing to light a printing material contg a photographic emulsion on a screen carried by a translucent detachable support, the emulsion being away from the source of light After development of the image, the print and the roller are transferred by inversion to a 2nd support, the translucent detachable support being detached before or after this transfer Cf. C. A. 25, 2653 and following abstrs

Color photography. SPICERS, LTD Fr. 693,220, June 30, 1930 A film of celluloid, viscose, cellulose nitrate, etc., for use in color photography is coated (1) with a layer of insulating material, such as a soln of cellulose acetate with or without plastifiers,

to prevent the penetration into the support, of the film of dye solns and agents for removing dyes, (2) with a layer of material compatible with the dyes and removers, such as cellulose nitrate, collodion or regenerated cellulose and (3) with a multicolor screen.

Color photography. STICKNEY, LEO. *Ir* 694258, July 1, 1930. Multicolored screens are applied to a film of cellulose acetate, nitrocellulose, etc., by applying on the surface of the film a constituent of coloring material and afterward printing successively on the surface of the film lines or other geometric designs, by means of one or more agents each containing a compound capable of reacting with the said constituent to form a differently colored dye.

Photographic developers. GUSTAV REDDFLEHN and WERNER MÜLLER (to Agfa Anso Corp.). *U S* 1,770,258, April 7. Photographic developers contain as a developing agent a *N*-monohydroxyethyl derivative of nuclear substitution products of 4-amino-1-hydroxybenzene are stable, very sol and capable of developing the latent photographic image in a Ag halide emulsion without formation of fog, even in solns free from KBr. Examples are given of developers containing 4-hydroxyethylamino-2-chloro-1-hydroxybenzene, 4-hydroxyethylamino-3-methyl-1-hydroxybenzene and 4-hydroxyethylamino-1-hydroxybenzene-2-carboxylic acid. *Cl C A*, 24, 3186, 25, 1749.

Developing light-sensitive layers with gases such as ammonia. I. MIL GRONAU (to Eugene Dickson Co.). *U S* 1,794,414, March 31. During development with a gas such as NH_3 , the paper being developed is slightly moistened with a liquid such as water containing the developing gas. App. is described.

6—INORGANIC CHEMISTRY

A. R. MIDDLETON

Advances in the preparation and utilization of a few rare elements. WALTER NODDACK. *Metallurgie* 21, (43)-4 671 2(1931).—A brief summary covering Ce, Hf, protactinium, W, I, Re and the Pt group. Re is obtained from molybdenite by soln in HNO_3 . Pptn. of $(\text{NH}_4)_2\text{PO}_4$, 12MoO_3 leaves Re in the filtrate, from which it is removed as the sulfide. The process, repeated many times, finally yields a 1% Re concentrate. This is oxidized and the oxide removed by sublimation. Reduction in H leaves the metal as a heavy, gray powder, m. slightly above 3400°Abs . Analytically Re belongs to the H_2S group, pptg. as Re_2S_7 from aq. solns of ReO_4^- . This ion is not readily reduced. B. A. SOLLE

Rhenium. IDA NODDACK and WALTER NODDACK. *Z anorg. Chem.* 44, 215-20 (1931), *Cl C A* 24, 3059.—A review covering the chem. properties of Re, its detection and data and the prepn. of HReO_4 and ReO_3 . To recover Re from residues, these are dissolved in dil. HNO_3 , boiled down and converted into chlorides by treating with HCl . The soln is evaporated to dryness, the salts are dissolved in water, the soln is made slightly alk. with KOH , treated with $(\text{NH}_4)_2\text{S}$ and then acidified with dil. H_2SO_4 . After standing 1 day the supernatant soln is again treated with $(\text{NH}_4)_2\text{S}$ and H_2SO_4 to ascertain whether the pptn. was complete. The ppt. is filtered, washed with 5% H_2SO_4 , then hot water, dried, washed with H_2O , then H_2O and dried again. The product is then pulverized and treated for 1 day with CS_2 , which dissolves the excess S. The residue is again pulverized and then added slowly to 4 parts of a 1:1 molten mixt. of KOH and KNO_3 in a Ni or Ag crucible. The cooked mass is dissolved in an excess of water and the soln is filtered. After concentrating and cooling the Re seps as a heavy crystalline powder of KReO_4 , very sparingly sol. in KOH . This is filtered, washed with a little water and 96% H_2O . The KOH soln. contains not more than 0.1 g. of KReO_4 per l. A test has shown that when the alk. soln. contains 100 times as much Cr, Mo, W, V, Nb, Ta, Ru, Os, Zn, Ga, Ge, As or Pb as Re, the first crystals yield KReO_4 , which contains not more than 0.2% of any of these elements. GEORGE CALINGAERT

Zirconium. II. Zirconium oxalate and diphenylidiazene zirconium. HOWARD S. GABLE. *J. Am. Chem. Soc.* 53, 3276-8(1931), *Cl C A* 24, 4996.—The use of MeOH solns. allows Zr compds. to be prepd. without danger of hydrolysis. Normal $\text{Zr}(\text{C}_2\text{O}_4)_2$ (I) has been prepd. for the 1st time by the interaction of solns. of ZrCl_4 and oxalic acid in MeOH . The white gelatinous ppt. is quite sol. in water, the basic salts of other workers being insol. Jefferson's work (*J. Am. Chem. Soc.* 24, 540, (1902)) on the pptn. of Zr salts by aniline was repeated with identical results in water or MeOH , the reaction being $2\text{C}_6\text{H}_5\text{NH}_2 + \text{ZrCl}_4 \rightarrow \text{C}_6\text{H}_5\text{N}=\text{N}-\text{C}_6\text{H}_5$ (II) + 4HCl . The solubilities of I and II were observed as follows: $\text{C}_6\text{H}_5\text{Cl}$, CS_2 , MeOH , EtOH , NH_4OH and HCl , both I and II insol.; $\text{C}_6\text{H}_5\text{NH}_2$, I slightly sol., II insol.; HCOH , I and II react,

H₂O, I very sol., II insol., AcOH, I slightly sol., II insol. Me₂CO, I slightly sol., II insol. HNO₃, I insol, II sol. ALBERT THOMAS FELLOWS

Synthesis of gas-metal compounds by sputtering. L. R. INGERSOLL. *J. Am. Chem. Soc.* 53, 2008-9(1931) —Ni sputtered in N under special conditions gives a nitride film, decomposed by heating at 750°, on heating to 150° in H₂, NH₃ and N₂ are formed. Compds. of Co and Fe with N and of Ni and H have also been prepd. C. J. WILST

Higher oxygen compounds of the eighth group of the periodic system. II. Compounds of nickel. D. K. GORALEVICH. *J. Russ. Phys.-Chem. Soc.* 62, 1165-88(1930), cf. *C. A.* 25, 654. —In the fusion of NiO with KNO₃ or with KClO₄ in the presence of alkali, there is formed aside from K₂NiO₄, a small amt. of a gray-green material which liberates O₂ on treatment with hot H₂O or with H₂SO₄ or HNO₃. This material is sol. in hot concd. HCl without decompos. but seps from the acid on cooling. Large yields of this material can be obtained by fusing in a silver cup a mixt. of NiO and KNO₃, and then adding to the hot fused mass K₂O, heating further, cooling, powdering and washing the product with ice water on a filter paper. The same reaction was repeated by fusing NiO with NaNO₃ and Na₂O, also with Ba(NO₃)₂ and BaO. The product of the reaction with Na compds. behaves in the same manner as the product with K compds. in regard to hot H₂O or HNO₃, H₂SO₄, or HCl. The product of the reaction with Ba compds. is not decomposed by hot H₂O or by heating to redness, but HNO₃ and H₂SO₄ liberate O₂ from it. When concd. HCl is added, Cl₂ is liberated on prolonged heating. Analysis of the fusion products showed them to be mixts. of compds. After repeated washing of the mixts. with dil. AcOH the residue remaining on the filter paper was a definite compd., e. g. in the case of the Ba fusion it was BaO₂NiO₂, in which Ni is octavalent. The pure BaNiO₄ is completely sol. in warm concd. HCl. On heating Na₂NiO₄ in concd. HCl there forms a green ppt. of Ni₂O₃ which is unstable and decomposes into NiO and O₂. K₂NiO₄ and BaNiO₄ behave similarly. On careful heating of Ni₂O₃ with concd. HCl a white unstable compd. NiO₄ forms, which decomposes in air into NiO and O₂. III. Nickel oxides. *Ibid.* 1577-1623. —Expts. were conducted to establish the existence and phys. and chem. properties of various Ni oxides, in regard to which there is confusion in the literature. (1) Oxides of Ni prepd. by electrolysis of powdered metallic Ni in strongly alk. Na pyrophosphate soln. contained Ni 67-73%, O 27-34%, no H₂O of crystn., active O by action of H₂O at 50° 6-16%, active O by action of H₂SO₄ 7-16%. Heating caused decompos. to NiO. (2) Hydrated NiO oxidized in alk. soln. in which Cl₂ was liberated electrolytically gave a product contg. Ni 66-73%, O 27-34%, active O by action of H₂O at 50° 5-16%, active O by action of H₂SO₄ 7-16%. (3) Ni salts, including chlorides, oxidized by means of K₂S₂O₈ in the presence of alkali gave products contg. H₂O 9-10%, Ni 64%, O 26%, active O in H₂O at 50° 0%, active O by action of H₂SO₄ 9% and Ni 60-73%, O 27-34%, active O by action of H₂O at 50° 7-16%, active O by action of H₂SO₄ 7-16%. (4) NiO oxidized by gentle heating with dry KNO₃ gave a product contg. Ni 67-72%, O 28-34%, active O by action of H₂O at 50° 8-15%, active O by action of H₂SO₄ 7-15%. Heating NiO with KClO₄ gave a product contg. Ni 66-70%, O 30-34%, active O in H₂O at 50° 11-16%, active O in H₂SO₄ 10-17%. (5) Heating Ni(NO₃)₂ or NiCO₃ gave products contg. Ni 71%, O 29%, active O by action of H₂SO₄ 9.5-10%. From these expts. the following oxides were established: Ni₂O, Ni₃O₂, Ni₄O₃, and NiO. In addn. some evidence was found for the existence of the following composite and less stable oxides: Ni₄O₂ = 4NiO₂NiO, Ni₄O₁₁ = 5NiO₂NiO, Ni₄O₈ = 2NiO₂NiO, Ni₄O₁₇ = 8NiO₂NiO, Ni₄O₁₃ = 5NiO₂2NiO, Ni₄O₁₆ = 7NiO₂NiO, Ni₄O₉ = 3NiO₂2NiO, Ni₄O₁₇ = 6NiO₂5NiO, Ni₄O₁₆ = 7NiO₂2NiO, Ni₄O₁₁ = 6NiO₂NiO, Ni₄O₈ = 2NiO₂3NiO, Ni₄O₇ = 2NiO₂3NiO, Ni₄O₄ = NiO₂NiO, Ni₃O₄·H₂O = NiO₂NiO·H₂O and NiO·Ni₂O₄ = NiO₂2NiO. Structural formulas for these compds. are suggested. S. L. MADORSKY

The preparation and structure of lower oxides of tungsten. J. A. M. VAN LIEMPT. *Rec. trav. chim.* 50, 347-6(1931). —An investigation to det. which oxides between W and WO₃ are stable was made by passing mixts. of CO and CO₂ and of H₂ and H₂O over WO₃ at fixed temps. Analysis of the products formed showed that of the lower oxides only WO₂(brown) and W₂O₃(blue violet) are stable. The latter is somewhat volatile. The heats of combustion of each of these oxides (W₂O₃ = 64,520 g. cal. per g. mol., W₂O₃ = 61,380) show that the structure is in fair agreement with the structural formulas proposed by Spitzin (*C. A.* 21, 544). M. McMAHON

The structure of ferric thiocyanate and the thiocyanate test for iron. H. I. SCHLESINGER AND H. B. VAN VALKENBURGH. *J. Am. Chem. Soc.* 53, 1212-6(1931). —In aq. solns. of Fe(CNS)₃, the color migrates to the anode during electrolysis. The absorption spectra of aq. solns. of Fe(CNS)₃, of aq. solns. contg. [Fe(CNS)₃]⁻ and of ether

sols. of anhyd $\text{Fe}(\text{CNS})_2$ are identical. In ether and benzene ferric thiocyanate has twice the mol wt. of the simple salt. On the basis of these results the authors conclude that the red color in the thiocyanate test for Fe is due to $[\text{Fe}(\text{CNS})_6]^{3-}$.

B. I. TIFFANY

Ammoniates of salts of the oxygen acids. M. A. RAKUXIN. *Bull. soc. chim.* [4], 49, 363-4 (1931).—By the action of dry NH_3 gas on dehydrated sulfates R has prep'd a no. of ammoniates: $\text{MnSO}_4 \cdot 4\text{NH}_3$ (brown), $\text{ZnSO}_4 \cdot 5\text{NH}_3$ (white), $\text{CuSO}_4 \cdot 5\text{NH}_3$ (blue), $\text{NiSO}_4 \cdot 6\text{NH}_3$, $\text{CoSO}_4 \cdot 6\text{NH}_3$ (white with a reddish tint), $\text{Cr}_2\text{SO}_4 \cdot 6\text{NH}_3$ (white), $\text{Ag}_2\text{SO}_4 \cdot 2\text{NH}_3$. Except with MnSO_4 and Ag_2SO_4 , the reactions are energetic. MgSO_4 does not react. With the exception of $\text{CuSO}_4 \cdot 5\text{NH}_3$, which forms $\text{CuSO}_4 \cdot 4\text{NH}_3 \cdot \text{H}_2\text{O}$, H_2O decomposes all of the ammoniates with the formation of the oxide or hydroxide of the metal. These ammoniates lose NH_3 when exposed to the air and more rapidly on heating. $\text{Ag}_2\text{SO}_4 \cdot 2\text{NH}_3$ on heating gives some $(\text{NH}_4)_2\text{SO}_4$ as well. Although AgNO_3 forms $\text{AgNO}_2 \cdot 6\text{NH}_3$, dry NaNO_3 and $\text{Ba}(\text{NO}_3)_2$ do not react with NH_3 . $\text{Cu}_2(\text{PO}_4)_2$ and $\text{K}_2\text{Cr}_2\text{O}_7$ likewise do not react.

W. C. LERNER

High-carbon carbides of the iron group. F. FISCHER AND H. BARR. *Ges. Abhandl. Akad. Wiss. Berlin* 8, 255-9 (1928).—When CO is passed at 270° over finely divided Co or CoO there is obtained a mass which appears to contain the carbide Co_3C . This substance is decomposed by H at 270° , half the C (presumably derived from the carbide Co_3C) being evolved as CH_4 . Similar addition of C to an equimol. mixt. of CuO and FeO affords the carbide Fe_3C in addn. to Fe carbonyl. It liberates CH_4 in proportion corresponding with the formula Fe_3C , formed apparently by separation of C from the unstable carbide Fe_3C_4 . Small quantities of hydrocarbons are also formed. The Fe, but not the Co, carbide affords, when treated with acids, a mixt. of hydrocarbons.

B. C. A.

The sulfides of carbon. KENNETH A. ROBE. *J. Chem. Education* 8, 867-74, (1931), cf. C. A. 25, 889.—K has reviewed the literature concerning these compds. and called attention to the similarity of the sulfides and the oxides of C. Unlike O, S exists on the surface of charcoal both in an adsorbed state and chemically combined. The phys. properties of COS are midway between those of CO_2 and CS_2 . Thirty-two references are given.

E. R. SCHEER

Phosphonates. II. Action of phosphine upon aluminum and beryllium halides. ROBERT HÖLTJE AND FRITZ MEYER. *Z. anorg. allgem. Chem.* 197, 93-102 (1931); cf. C. A. 24, 3721.—Mol. compds. of PH_3 are quite similar to the corresponding thiohydrides, and salts which add on H_2S are also capable of taking up PH_3 . By heating AlCl_3 and AlBr_3 with PH_3 at 70° , $\text{AlCl}_3 \cdot \text{PH}_3$, m. $81-3^\circ$, and $\text{AlBr}_3 \cdot \text{PH}_3$, m. $514-8^\circ$, were formed. In the case of AlCl_3 it was necessary to raise the temp. to $120-30^\circ$ to obtain $\text{AlI}_3 \cdot \text{PH}_3$, m. $148-50^\circ$. In no case was the amt. of PH_3 taken up increased by increasing the pressure. Similar compds. with PH_3 were not obtained in the case of Be halides. The phosphonates mentioned above are white cryst. compds. which readily sublime and are decomposed by water to give PH_3 and the Al salt. These compds. are readily fused, and in fusing they deposit a small quantity of a yellowish red solid assumed to be P_2H_4 . Vapor pressure detns. showed that PH_3 was firmly held in these compds., this high degree of stability being attributed to the great deformability of PH_3 .

H. STOEFTZ

Calcium oxalate from calcium cyanamide. GEORGE BARSKY AND G. H. BUCHANAN. *J. Am. Chem. Soc.* 53, 1270-6 (1931).—Autoclaving $\text{Ca}(\text{CN})_2$ and CaCN_2 with water gives CaCO_3 , CaC_2O_4 , and $\text{Ca}(\text{COOH})_2$, and if the H_2O is replaced by $\text{Ca}(\text{COOH})_2$, solns. increased yields of oxalate are obtained. Although the yield of oxalate from cyanide is low, the process has com. value because small amts. of cyanide can be added to cyanamide which is being autoclaved for NH_3 , as all the N of both compds. is liberated as NH_3 , and a part of the C of both compds. is converted to valuable by products. Simple treatment of the ordinarily discarded autoclave solids yields CaC_2O_4 . The process is protected by U. S. P. 1,717,353, (Am. Cyanamide Co.; C. A. 23, 3933). The reaction taking place is represented $2\text{CaCN}_2 + \text{Ca}(\text{CN})_2 + 10\text{H}_2\text{O} = 6\text{NH}_3 + 2\text{CaC}_2\text{O}_4 + \text{Ca}(\text{OH})_2$, the mechanism being a sort of mixed polymerization, H_2NCN and HCN reacting with each other, the reaction velocity being dependent upon the H^+ ion concn. of the soln., just as cyanamide reacts with itself to give dicyanamide (cf. C. A. 24, 1085) and HCN reacts to give azulene compds. This reaction is thus represented $\text{H}_2\text{NCN} + \text{HCN} \rightarrow \text{H}_2\text{NCNH-CN}$ (hypothetical), $\text{H}_2\text{NCNH-CN} + 4\text{H}_2\text{O} \rightarrow (\text{NH}_4)_2\text{C}_2\text{O}_4 + \text{NH}_3$.

ALBERT THOMAS FELLOWS

Copper quadratoxide. M. J. MURRAY. *J. Phys. Chem.* 35, 1011-24 (1931).—By a study of the rate of reduction of CuO in H_2 , it is shown that Cu quadratoxide

does not exist. Conclusion The olive-green ppt. reported in the literature is probably a mixt. of metallic Cu and hydrous or anhyd. Cu_2O B. E. TIFFANY

Anhydrous copper halide carbonyls. OTTO H. WAGNER *Z. anorg. allgem. Chem.* 196, 364-73 (1931).—The reaction between solid cuprous halides and CO was studied at elevated pressures and the existence of $\text{Cu}_2\text{Cl}_2 \cdot 2\text{CO}$ and $\text{Cu}_2\text{Br}_2 \cdot 2\text{CO}$ was established. In the case of the iodide, the reaction was so sluggish that less than 1 mol. of CO was taken up per atom of Cu. The absorption of CO occurred even in the complete absence of water, although water does exert a catalytic effect upon the reaction velocity.

H. STOBERTZ

Iron titanates. B. PESCH *Gazz. chim. ital.* 61, 107-11 (1931).—A survey of the literature shows that FeTiO_3 (ilmenite) is the only Fe titanate the prepn., compn., and properties of which are definitely known. A systematic study was therefore begun of Fe titanates, and as a beginning FeTiO_3 was prepd. in a new way. A mixt. of Fe_2O_3 (free of Fe_3O_4), TiO_2 (calcd. quantity) and NaCl (3 times as much) was heated 15 hr. at 900° in a current of CO , and cooled in the absence of air. The product was digested in water, treated with very dil. HCl and washed until there was no longer a reaction for Cl. A black powder of the compn. FeTiO_3 was obtained. The method of Rossi for the detn. of Ti (pptn. by boiling of H_2TiO_4 in a current of SO_2) was used. The FeTiO_3 was dissolved by treatment with H_2SO_4 . Under the same conditions but with proportions of Fe_2O_3 and TiO_2 corresponding to $2\text{FeO} \cdot \text{TiO}_2$, a black powder was obtained, which analysis showed to be ferrous orthotitanate, Fe_2TiO_4 . Likewise TiO_2 (2 mols.), Fe_2O_3 (3 mols.) and NaCl heated at 1000° and washed free of NaCl yielded ferric orthotitanate, $\text{Fe}_2(\text{TiO}_4)_3$, dark maroon red, birefringent, yellowish by transmitted light. Attempts to prep. $\text{Fe}_3(\text{TiO}_4)_2$ under the same conditions gave in all cases mixts. of $\text{Fe}_2(\text{TiO}_4)_3$ and TiO_2 . Fe_2TiO_4 was described by Doss (*Z. Kristallographie* 1892, 566) and its structure confirmed by Pauling (*C. A.* 24, 2093). By using for its prepn. the method with $\text{Fe}_2(\text{TiO}_4)_3$ (*loc. cit.*), the results obtained were neg. The results in general show that of all the titanates, $\text{Fe}_2(\text{TiO}_4)_3$ is the most easily formed. The compd. obtained by Hautefeuille (*Ann. chim. phys.* 4, 172 (1864)) was probably $\text{Fe}_2(\text{TiO}_4)_3$. By following the procedure used by König and von der Pfordten (*Ber.* 22, 1405 (1889)) only complex mixts. contg. chiefly Fe_2O_3 were obtained. X-ray examn. (powder method) of $\text{Fe}_2(\text{TiO}_4)_3$, FeTiO_3 and Fe_3TiO_4 showed that all 3 are definite compds. rather than mixts. The roentgenographs are shown in simplified form.

C. C. DAVIS

Ferrous ferrite. L. Ferrous ferrite from ortho- and meta-ferric hydroxide, ALFONS KRAUSE AND J. TULECKI *Z. anorg. allgem. Chem.* 195, 228-38 (1931).—Ferrous ferrite was made at 18° from metaferic hydroxide, dispersed in 0.01 N HCl, and FeCl_3 soln. by the addn. of NH_4OH (I) or from FeCl_3 and FeCl_2 solns. with excess NH_4OH (II). Two types of compds. result, (I) being blue black in color, stable in the air and approaching $\text{Fe}(\text{FeO}_2)_2$ in compn. ($\text{FeO} \cdot \text{Fe}_2\text{O}_3 \cdot \text{H}_2\text{O}$, 1:1:3.051), while (II) varied considerably in compn. in consequence of oxidation, contained more water, and was brownish black. All ferrous ferrites prepd. were magnetic.

L. P. HALL

Potassium chlororhenate, E. ENK *Ber.* 64B, 791-7 (1931).— K_2ReCl_6 is prepd. according to the reaction $\text{KReO}_4 + 3\text{KI} + 8\text{HCl} = \text{K}_2\text{ReCl}_6 + 2\text{KCl} + 3\text{I} + 4\text{H}_2\text{O}$. This salt is analogous to others of the type $\text{M}_2'\text{M}''\text{X}_6$. It crystd. in yellow-green octahedra with the following physical consts: $d_{20} 3.34$, sp. cond. at 25° , $N/32$ soln., $3.47 \times 10^{-3} \Omega^{-1}$, equiv. cond. 119.7 Ω^{-1} . It is but slightly sol. in water. It dissociates into 2K^+ and ReCl_6^- .

MARY I. LEAR

The reduction of potassium perrhenate. HENRY V. A. BRISCOE, PERCY L. ROBINSON AND ERIC M. STODDART *J. Chem. Soc.* 1931, 666-9.—The reduction of aq. KReO_4 solns. by Zn, hydrazine hydrate and other reducing agents is reinvestigated, previous workers gave conflicting results. It was shown that the first product of reduction of KReO_4 with Zn or hydrazine hydrate is a yellow colloidal soln., most probably of hydrated ReO_2 . From these solns. the black $\text{ReO}_2 \cdot 2\text{H}_2\text{O}$ is then pptd. This material can be quantitatively pptd. and, also, can be dehydrated without decomposition. The charge on the colloid is shown to be pos.

J. BALOZIAN

Formation of sodium hydrosulfide from sodium sulfide. M. O. KHARMANDARYAN AND K. I. BRODOVICH *Zhurn. Prikladnoi Khim.* 3, 1023-30 (1930).—For Na_2S (in water soln.) + $\text{H}_2\text{S} \rightleftharpoons 2\text{NaHS}$ the const. of the reaction is 0.0126. The max. yield of NaHS is 89.7%. Adsorption of H_2S by cryst. Na_2S yields pure NaHS in soln., but its crystn. is difficult. Adsorption of H_2S by dried Na_2S yields dry NaHS. The reaction is speeded up if Na_2S is wetted with NaHS soln.

V. KALICHEVSKY

Reactions with sodium nitroprusside. TITO PAVOLINI. *Boll. chim.-farm.* 69,

713-6, 719-22(1930) —P reviews the reactions given by Na nitroprusside (I) + alkali with the following groups of compds: alkali sulfides, sulfites, derivs of NH_3 , NH_2 and H_2NOH , aldehydes and ketones, hydantoin and derivs, pyrrrole and indole and some cyclic dinitro derivs. He reports original work with compds of these and some other groups. The blue color given by alkali hydrosulfides is caused by colloidal Berlin blue, which is formed by the interaction of I with H_2S . Na_2Se gives a deep blue green color with I, the soln decolorizes rapidly and a red ppt of amorphous Se appears. With Na_2Te black, amorphous Te is pptd. The Bodecke reaction is not only very sensitive but is specific for sulfites. The aliphatic α diketones give pos reactions, the liquids rapidly turn yellow since in the presence of alkali the α diketones are converted into p-quinones. With β - and γ -diketones reddish violet colors which then turn to bright red were observed. Paraldehyde, metaldehyde, chloral, glucose and fructose do not react with I. Carvone, pulegone, ionone and citral gave red colors of various hues, while the results with camphor, lenchone, menthone and citronellal were neg. This difference in behavior is explained by the assumption that double bonds in the close vicinity of the CO group favor the occurrence of the reaction Me_2CO in the presence of I and piperazine gives a rose violet color, with piperidine and I there appears an intense red color, which becomes purple with AcOH . This reaction allows the differentiation between Me_2CO and AcEt , because the latter gives an intense dark blue color with I and piperidine. Expts with compds with a structure similar to that of hydantoin showed that pos reactions are given by uracil, 4-methyluracil, 5-methyluracil and allantoin. Barbituric acid and veronal gave neg results. Neg results were obtained with α -methylindole, β -methylindole, oxindole, dioxindole, carbazole, acridine, diphenylamine, pyridine, quinoline, benzimidazole, furan and thiophene. On the basis of these

results P concludes that the reaction is characteristic for the grouping $\text{—C} \begin{array}{c} \text{NH}_2 \\ \text{CH} \end{array} \text{CH}$

Some cyclic dinitro derivs with the NO_2 groups preferably in the m position, e.g., m -(O_2N) $_2\text{C}_6\text{H}_3$, react with I and alkali. o -, m - and p -(O_2N) $_2\text{C}_6\text{H}_3\text{OH}$, (O_2N) $_2\text{C}_6\text{H}_3\text{OH}$ and o -(O_2N) $_2\text{C}_6\text{H}_3\text{CHO}$ do not react. (O_2N) $_2\text{C}_6\text{H}_3\text{Me}$ in alc soln gave a rose-violet color on addn of KOH alone. Thiourea, diphenylthiourea, o -phenylenethiourea and to a less degree also thiosinamine in alc soln give red violet colors with I and alkali. This reaction is due to the decarboxylation of the compds by the addn of alkali with the subsequent formation of alkali sulfide. The reaction is neg with KCNS , NH_2CNS and Li_2S . Some polyhydric phenols give color reactions with I and alkali. Intense green colors were obtained with pyrocatechol and resorcinol, while green colors of various hues, generally with a tendency toward brown, were observed with hydroquinone, pyrogallol, phloroglucinol and hydroxyhydroquinone; the color changed to a bluish green on addn of AcOH . The reactivity of I is especially due to the tendency of Fe to become bivalent, and of the coordinated nitroso group to convert itself into a nitrito group by the absorption of O. Thus in the case of the formation of Na nitropentacyanide the reaction takes the course $\text{Na}_2\text{Fe}(\text{CN})_5\text{NO} + 2\text{NaOH} = \text{H}_2\text{O} + \text{Na}_2\text{Fe}(\text{CN})_5(\text{NO}_2)$. This latter compd, which crystallizes with 10 mol H_2O , has a certain analogy with the alkali ferrocyanides, even with respect to the color. The bright red compd obtained in the Bodecke reaction was analyzed and its compn was established as $\text{Zn}_2\text{Fe}(\text{CN})_5\text{NO} \cdot \text{SO}_4$. The reaction mechanism between I and aldehydes and ketones was studied in particular. A pos color reaction requires at least the presence of the structure $>\text{CHCO}$ and occurs as a rule with compds with short chains or with a double bond in the immediate vicinity of the CO group. From a mixt of I + NaOH in concd aq soln and Me_2CO or AcEt in abs alc, P succeeded in pptg the colored compds by addn of an excess of alc. A pos reaction was obtained with I and ether which had been exposed to light, this is due to the presence of a small quantity of H_2CCHOH , which is formed by photooxidation of the ether. In general there must exist a certain relation between the compds giving a pos reaction with I and the facility with which they are transformed into the enol form, e.g., $\text{AcEt} \rightarrow \text{H}_2\text{CCHOH}$, $\text{Me}_2\text{CO} \rightarrow \text{Me}(\text{HO})\text{CCH}_2$. The same color reactions observed with I, alkali and AcEt or Me_2CO , resp, were also obtained by treating Na nitropentacyanide with AcEt or Me_2CO , resp, in alc soln, the reaction with Me_2CO is rather slow. Both reactions are accelerated by the addn of a small quantity of alkali. G. SCHWOCH

Nitrogenation of chromium. L. DUPARC, P. WENGER AND W. SCHUSSELE. *Helv Chim Acta* 13, 917-29(1930).—By following the same procedure used in the prepn of the nitrides of Mn (C.A. 23, 4504), the nitrides of Cr were investigated, 2 different samples of Cr were used. The absorption threshold, optimum temp and max amt. of N fixed varied with the Cr employed. The absorption thresholds for the 2 samples

examd. were 650° and 750°. The amt of N fixed was independent of time but increased with pressure. Lu_2N used as a catalyst lowered the absorption threshold but appeared to diminish the fixation of N, possibly on account of the formation of an easily dissociable Cr-Li-N complex. C. H. PRET

The action of sulfuric acid on copper. F. DE CHAVES. *Anales soc. españ. fis. quim.* 29, 177-81 (1931).—Pure, concd H_2SO_4 , free from As , was caused to react with electrolytic Cu at 55° for 6 hrs. The residue contained anhyd CuSO_4 , S and a deposit contg S 8.5 and Cu 69.7, the balance being O or O and H. The deposit does not correspond to the Maumene oxysulfides or to any known Cu ores. E. M. S.

The decomposition of carbonyl chloride by heat. II. ALFRED STOCK, WERNER WUSTROW, HERMANN LUX AND HANS RAMSER. *Z. anorg. allgem. Chem.* 195, 140-8 (1931), cf C. A. 19, 3228.— COCl_2 was heated at 400° in a glass tube with and without catalysts. Decompn to CO and Cl_2 , detd from the resulting CO content, was 23.6% in 143 hrs in glass alone, 23.6% on AlCl_3 in 90 hrs, 23.3% on SiO_2 in 115 hrs. No decompn to CO_2 and CCl_4 was observed in these cases. On activated charcoal the decompn CO was 22.7% in 88 hrs. CCl_4 was formed by a secondary reaction of C and Cl. An attempt to produce COCl_2 from CO_2 and CCl_4 at 400° to 500° was unsuccessful. L. P. HALL

The thermal decomposition of silver subfluoride. R. SCHOLDER AND K. TRAUTSEN. *Z. anorg. allgem. Chem.* 197, 57-61 (1931).—An improved method of manufacturing Ag_2F is given as follows. Add silver carbonate to a hot 40% HF soln until no more dissolves. Add 2 g of NH_4F and filter in the dark. Electrolyze the filtrate in a 100 cc Pt dish at 50° with a silver anode, 1.4 v tension and 0.07-0.10 amps. In 48 hours 20 g of pure cryst Ag_2F is obtained. Dry the crystals on filter paper. Ag_2F decomposes into Ag and AgF between 100° and 200°, no F is evolved, even at 700°. A. L. HENNE

The reduction of nitric oxide. M. L. NICHOLS AND C. W. MORSE. *J. Phys. Chem.* 35, 1239-52 (1931).—NO dissolves in water to form nitrohydroxylamic acid ($\text{H}_2\text{N}_2\text{O}_3$) or a compd of the same constitution. The authors studied the reduction of NO in soln by SnCl_2 and TiCl_3 . $\text{H}_2\text{N}_2\text{O}_3 \rightarrow \text{H}_2\text{N}_2\text{O}_2 \rightarrow \text{H}_2\text{N}_2\text{O} \rightarrow \text{NH}_2\text{OH} \rightarrow \text{NH}_3$. At temps above 0° the reduction by SnCl_2 depends on the rate of formation of $\text{H}_2\text{N}_2\text{O}_3$. At temps below 25° the rate of reduction by TiCl_3 depends on the rate of the reducing reaction, while at temps above 25° the rate of formation of $\text{H}_2\text{N}_2\text{O}_3$ is the limiting factor. NH_2OH is apparently the end point of the SnCl_2 reaction. Some observations were made on the catalytic effects of reagents or their oxidation products. B. E. TIFFANY

A complex cyanide of quadrivalent vanadium. AL. YAKIMACH. *Compt. rend.* 191, 789-90 (1930).—Addn of concd $\text{V}(\text{OAc})_4$ to KCN of the same concn, with stirring, causes the soln to become green, and after several weeks, green tetragonal prisms are obtained. The crystals decompose rapidly; so they are washed with EtOH. An excess of $\text{V}(\text{OAc})_4$ destroys the compd, as does an excess of the KCN (albeit more slowly). The crystals decompose at 150° with blackening; they also decompose on soln in water. Detns. of K, V and (CN) $_4$ gave the formula $\text{K}_2\text{V}(\text{CN})_4$. I. J. PATTON

7—ANALYTICAL CHEMISTRY

W. T. HALL

Hydroxylamine as a precipitating agent. JOSÉ CASARES ROLDÁN. *Anales soc. españ. fis. quim.* 29, 158-61 (1931). cf C. A. 25, 261.— NH_2OH reacts with $\text{Cu-Fe}(\text{CN})_4$ to discharge the color and give a beautiful fluorescence. A complex of $\text{K}_4\text{Fe}(\text{CN})_6$, NH_2OH and NH_3 with Cu salts is being investigated. E. M. SYMES

Studies concerning the relations between reaction sensitivity and molecular size in the case of organic reagents. JOSEF V. TANCIVNA. *Mikrochemie* 3, 229-41 (1931).—The tests for Cu and for molybdate with xanthogenates and the test for Ag with derivs of rhodanic acid were studied. It was found that by increasing the mol wt., the soly of the resulting compds. was diminished so that the test became more sensitive. In the tests with various rhodanates there was also a marked deepening of the color which also made the test much more sensitive. The general conclusion is drawn that increase of mol wt. increases the sensitiveness and the introduction of the benzene ring increases the depth of color. W. T. H.

Systematical qualitative analysis by means of modern drop reactions. C. J. VAN NIEUWENBURG. *Mikrochemie* 3, 199-219 (1931).—A method is given of systematic

examn. by means of drop tests suitable for the detection of about 0.5% of 25 cations and 18 anions. not over 0.2 g. of substance is used for the complete analysis. Attempts to avoid the use of H_2S as reagent proved unsuccessful, in the sepn. of the sulfides pptd. in acid soln. the use of KOH is preferred to alk. polysulfide. Another new feature is the pptn. of Al and Zn from alk. solns. by neutralizing to pH 8.5 with thymol blue as indicator. The advantages gained by the use of the centrifuge in place of ordinary filtration are also emphasized. The procedure in the hands of students has shown that the time required for a complete qual. analysis is at least halved. The following tests are recommended: 1 or Ag the fresh ppt. of AgCl is treated with a 0.03% soln. of rhodanine in acetone, a red color results. Hg gives the same test but cannot be present when the chloride ppt. has been properly treated. a little KCN prevents interference by Hg . 1 or Hg the HgS ppt. is dissolved on the spot plate in a drop of Hr. aq. , the excess Hr. is removed by treatment with H_2SO_4 and phenol and a small drop of 1% diphenylcarbazide soln. in alc. is added, a blue color results. For Pb the hydroxide or sulfate is treated on the spot plate with NaOH and H_2O_2 , the supernatant liquid is removed from the dark brown PbO_2 by filter paper, after which NH_4Cl , a drop of dil. AcOH and a drop of 0.5% tetramethyl diaminodiphenylmethane soln. in alc. + 5% AcOH is added, a bright blue color results. For Bi the alk. stannite test is used in the presence of a little $\text{Pb}(\text{OAc})_2$, which accelerates the reduction to metallic Bi . 1 or Cu , to the soln. on the spot plate a drop of 10% ZnSO_4 soln. is added and a white ppt. (bluish violet in the presence of Cu) of Zn Hg thiocyanate formed by adding a soln. 30 g. HgCl_2 and 33 g. NH_4CNS in 100 cc. of water. Cd is detected by heating in a closed tube with soda, collecting the sublimate in a restriction of the tubing and converting it into CdS by passing S vapors over it. As is detected by the Gutzeit test and Si by the deposit formed on Sn . Sn is proved present by reducing the HCl soln. with Al and adding a little 0.25% soln. of cacotheline, a brown violet color results. 1 or Al , the tests with alizarine S and with the NH_4 salt of aurintricarboxylic acid are both recommended. 1 or Fe , the thiocyanate test is used. Cr is detected by converting the hydroxide to chromate with KOH and Hr. and adding a 1% soln. in alc. of diphenylcarbazide after removing excess Br with phenol. U is detected by the $\text{K}_2\text{Fe}(\text{CN})_6$ test and Co by the reaction of Vogel or by the von Knorre test with α -nitroso- β -naphthol. For Ni the familiar glyoxime test is used and for Mn the red color formed in AcOH soln. with a 1% soln. of benzidine acetate in dil. AcOH . The formation of MnO_4^- by the persulfate reaction in the presence of Ag^+ is also recommended. Zn is detected by adding a trace of Cu^{++} and testing with the $\text{NH}_4\text{Hg thiocyanate}$ soln. used in the test for Cu . Mg is detected by boiling an alk. suspension of $\text{Mg}(\text{OH})_2$ with a little titan yellow. For Ba , Ca , Sr , Na , K and NH_4 , the usual tests are recommended. The tests for the anions involve well known procedures, but particular attention is paid to distinguishing between nitrite and nitrate, formate and acetate and the various forms in which S can exist. Blow pipe tests, flame tests and closed tube tests are used to an extent that will please all the older chemists. The procedure for making a complete qual. examn. is given in detail.

W. T. H.

New method to decompose rocks containing sulfur. T. GEDON, *Magyar Chem. Folyóirat* 37, 71-4 (1911).—Mix 1 g. of the sample with 6-8 g. of a mixt. contg. equal parts of Na CO_3 , KMnO_4 and MgO , cover with 3 g. of the above mixt. and heat the crucible gently on an asbestos plate for 15 min. Then apply the full flame of a Bunsen burner for 6-8 min. Cool, leach out the contents of the crucible with hot water and reduce any MnO_4^- by adding a few drops alc. Filter, make acid with HCl and det. SO_4^{--} by pptn. with BaCl_2 as usual. The method was tested with bauxite and with coal, the results agreed with those obtained by 7 other standard methods.

S. S. DE FINÁLY

Use of grid electron valves for potentiometric titration. G. HATOS, *Kísérlet Kozlemények* 31, 299-307 (1928).—The method of Treadwell was used with the help of a Philips A 409 valve with good results. *Ibid* 32, 654-62 (1929).—Results obtained with the method of Treadwell and Kohlrausch were quite as exact as those obtained by the app. of Kohlrausch.

S. S. DE FINÁLY

Telephone as zero instrument in the electrometric compensating method. G. HATOS, *Kísérlet Kozlemények* 33, 17-9 (1930).—A 4600 ohm telephone was used with 1000-mm app. Data agreed to 0.1 μm in case the inner resistance of the app. was not larger than 12 000 ohms.

S. S. DE FINÁLY

Electroanalytical separations in ammoniacal fluoride solutions. I. Separation of copper from arsenic and antimony. N. HOWELL FURMAN, *Ind. Eng. Chem., Anal. Ed.* 3, 217-8 (1931).—Dissolve the mixt. of Cu , As and Sb in 3 to 5 cc. of 48% HF and 25 cc. of HNO_3 (1 vol. of acid of d. 1.42 and 4 vols. H_2O). Add 1 to 2 g. of $\text{K}_2\text{S}_2\text{O}_8$. Boil the acid soln. 2 to 3 min. and neutralize immediately with NH_3 . Sb and As are thus

brought to a quinquevalent state and will not be reduced during the electrolysis of the cold, strongly ammoniacal solution the correct amount of Cu is deposited A. L. H.

Microchemical determination of nitrogen. I. FLEISCHNER. *Kisérlet Kémiények* 32, 113-22(1929) —Description of present methods and data. S. S. DE F.

Colorimetric determination of phosphorus. LÁSZLÓ URBANEK. *Mezőgazdasági Kutások* 4, 39-59(1931) —Factors influencing the detn. of P by the blue color of reduced phosphomolybdic acid are discussed The development of the color depends also on the pH of the soln and on the duration of the reaction Generally it takes 2 hrs to reach the final color Oxalic and citric acids retard the reaction materially, HNO_3 , HCl , H_2SO_4 and CCl_3CO_2H have little effect and $AcOH$ has practically none Salts do not affect the reaction except by their buffer effect The influence of reducing materials is annulled in 2 hrs Z. I. KERTESZ

Determination of iodine in inorganic material by microchemical methods. GULBRAND LUNDE. *Mikrochemie* 7, 337-66(1929) —A description is given of the many papers published on this subject from 1876 to the present time W. T. H.

Influence of nitrites, thiocyanates and some organic substances on the iodine-starch reaction. Z. ERNST. *Biochem. Z.* 232, 346-51(1931). —The acid nitrite-starch reaction for detecting I is much more sensitive with saliva than with pure aq. solns because of the presence of $KSCN$ in the saliva The I-starch reaction in an acid soln is inhibited by nitrite, and the inhibiting action is removed by thiocyanates Uric acid, peptone and to a lesser degree also leucine, trypsin and serum albumin act like the thiocyanates S. MORGULIS

A color test for cobalt. M. GUTIÉRREZ DE CELIO. *Anales soc. españ. fis. quim.* 29, 262-3(1931) —The blue color resulting from the formation of a double thiosulfate of Co and Na can serve as a sensitive test for Co As little as 2.3×10^{-4} mg. of Cr can be detected E. M. SYMMES

Reactions of the nitroso derivative of R salt with various inorganic salts. A. BERNARDI AND M. A. SCHWARTZ. *Ann. chim. applicata* 21, 45-50(1931) —R salt has been proposed as a reagent in the identification of Co (van Klooster, *C. A.* 15, 1671). It has been found that the following compds. of Pb, Ag, Ba and Ca are also obtained from aq. solns of R salt on adding the nitrates of the first two, and Ba and Ca as chlorides: $(C_{11}H_7O_4NSNa)_2Pb$, $Pb(NO_3)_2$ (red), $(C_{11}H_7O_4NSNa)_3Ag$, $AgNO_3 \cdot 3H_2O$ (lemon-yellow), $(C_{11}H_7O_4NSNa)_2Ba$, $BaCl_2 \cdot 6H_2O$ (orange), and $(C_{11}H_7O_4NSNa)_2Ca$, $CaCl_2$ (green) The Ba salt is insol. in HCl solns as well as $EtOH$ solns, while the Ca salt is sol. in HCl , but insol. in alk. (NH_4OH) $EtOH$ soln. In this way these 2 elements can be sep'd A. W. CONTIERI

Determination of manganese in ferrosilicons by the volumetric method. J. H. D. BRADSHAW. *Foundry Trade J.* 44, 311(1931) —Fe-Si alloys with not over 18% of Si can be attacked satisfactorily by a mixt. of 30 cc. HCl , 15 cc. HNO_3 and 5 cc. of H_2SO_4 . After fuming, to remove HCl and dehydrate SiO_2 , the silica can be det'd and the persulfate method for Mn applied to the filtrate from the SiO_2 . With higher Si contents, the original material must be fused with alkali carbonate at the start W. T. H.

Analyzing chromium solutions. I. L. NEWELL AND W. H. KEEFE. *Metal Ind.* (N. Y.) 29, 206-7(1931) —A complete scheme of analysis is given for analyzing plating solns. for CrO_4^{--} , Cr^{+++} , SO_4^{--} , Cu^{++} , Fe^{+++} , Ni^{++} and Zn^{++} . The CrO_4^{--} is det'd by taking 1 ml. of the soln, adding 400 ml. of water, 5 ml. of 85% H_3PO_4 and 0.6-0.8 ml. of a soln obtained by dissolving 0.1 g. of diphenylbenzidine in 10 ml. of concd. H_2SO_4 and 90 ml. of $AcOH$. After a few min. add 3 g. of $NaOAc$ and titrate with 0.2 N soln. of Mohr's salt to the disappearance of the blue color To det. Cr^{+++} , dil 25 ml. of the soln to 200 ml. and ppt with NH_4OH . Dissolve the $Cr(OH)_3$ ppt in H_2SO_4 and oxidize to CrO_4^{--} by adding Na_2O_2 . After boiling to decompose excess Na_2O_2 , treat as described above for detg. CrO_4^{--} . To det. SO_4^{--} , dil 10 ml. of soln with 20 ml. of water, add 7 ml. of HCl , 25 ml. of $AcOH$ and 15 ml. of alc. Boil 30 min. and add 10 ml. of $BaCl_2$ reagent To det. Cu , take 10 ml. of soln, dil. to 200 ml., add an excess of $Pb(NO_3)_2$ soln and filter off the $PbCrO_4$. To the filtrate add an excess of Na_2O_2 and filter off insol. hydroxides of Cu, Fe, Ni, etc., using the filtrate for the Zn detn. Dissolve the ppt in HCl and H_2SO_4 , sat. with H_2S and filter off the CuS ppt. Dissolve this in HNO_3 and det. Cu electrolytically Det. Ni in the filtrate from CuS by the dimethylglyoxime method after first pptg. Fe as $Fe(OH)_3$ and weighing as Fe_2O_3 . In this final filtrate, combined with the filtrate from the Na_2O_2 treatment in the detn. of Cu , det. Zn as phosphate W. T. H.

The determination of germanium. W. GEILMANN AND K. BRÜNGER. *Z. anorg. allgem. Chem.* 196, 312-20(1931) —The smallest amt. of Ge which could still be detected spectroscopically was 0.25 γ in 0.025 cc. of soln. The C electrodes (boiled in HCl) used

in obtaining spark spectra at 12,000 v. a. c. were superior to Au and Pt electrodes, which decreased the sensitivity 75 times. A table shows the changes in the Ge spectrum with concn. (1.0-0.001) g. of Ge in (100 cc.) Ge could be concd. by pptg. it in 3-4 N HCl soln. with H₂S and redissolving the ppt. in N KOH. When As, Se and Sb were present, Ge was sepd. by distn. from 3-4 N HCl in a current of Cl₂ into 3-4 N HCl at 0°. Ge in the fitter soln. was concd. by pptn. with H₂S in the presence of Na₂SO₄. The colloidal S formed brought about better pptn. of GeS₂. Suitable methods for the detn. of Ge in oxides, silicates, cassiterite, sulfide ores and metals are outlined. **IRANK URBAN**

Tests for sodium with the Streng-Kolthoff reagent. **R. MONTEGUT AND R. DA SIRADA. *Anales de quim. fis. quim.* 29, 255-61 (1931)**—The Streng-Kolthoff reagent is (1) 1.0 (C₂H₅O)₂, 10-20% AcOH, 6, H₂O to make 50 cc. and (2) Zn(C₂H₅O)₂, 20, 20% AcOH, 3, H₂O to make 50 cc. 1 and 2 being mixed hot and filtered after standing 24 hrs. To det. Na in a sol. salt dissolve 0.15 g. of sample in 0.5 cc. of H₂O, add 1.5-2.0 cc. of Zn(C₂H₅O)₂ soln., stir and filter. This serves to remove K⁺, which is likely to ppt. with Na when much K is present. Add 2 cc. of the Streng-Kolthoff reagent to 0.5 cc. of the filtrate. Turbidity of pptn. begins in a few min. Microscopic examn. shows typical crystal formation. Detns. are accurate on 0.15 g. KCl in 0.2 cc. of 0.01 N NaCl. **F. M. SYMMES**

Use of phenolic acids in the detection, separation and estimation of metals. V. Separation of copper from cadmium, and their subsequent estimation. **PARVATHI NATH DAS GUPTA AND HARIBOLA SAMA. *J. Indian Chem. Soc.* 8, 10-21 (1931)**—By adding a 1% soln. of gallic acid to the boiling soln. of Cu and Cd in the presence of alk. acetate, a voluminous brown ppt. is formed which can be filtered off and ignited to CuO. The ppt. is bulky and can serve for the detn. of small quantities of Cu. The Cd is not pptd. by gallic acid under these conditions. **W. T. H.**

Delicate microchemical reaction of copper salts and certain of the other heavy metals. **I. M. KORENMAN. *Pharm. Zentralhalle* 72, 225-6 (1931)**—The delicate test for pierce acid by pptn. with Cu salts in the presence of NH₃ as suggested by Rymasz is reversed in testing for Cu, Ag, Co, Ni, Cd, Hg, Au and Pb. The reagent consists of 2 parts of a satd. aq. soln. of CaH₂(NO₃)₂OH to 1 part 10% NH₃. This reagent produces in highly diltd. solns. of Cu salt a bright greenish yellow ppt. of characteristic type. The test is delicate to 0.05 μg. Cu (1 μg. = 0.00001 g. = 1 microgram). With Ag, Co, Ni, Cd, Hg, Au the test is sensitive to 20 μg., 0.7 μg., 0.1 μg., 1 μg., 0.15 μg. and 0.2 μg. resp. **W. O. F.**

Quinoline as a microchemical reagent for certain heavy metals. **I. M. KORENMAN. *Mikrochemie* 3, 223-8 (1931)**—See *C. A.* 25, 557. **W. T. H.**

Rapid industrial methods for the analysis of metallurgical products. **ROGER ARZENAU. *Bull. ass. tech. fondre* 5, 31-8 (1931)**—A general review is given of rapid methods for plant control in the analysis of steel for S, Mn, P, S and C. The advantages of using factor weights and reagents in the form of compressed lozenges are emphasized. **W. T. H.**

Determination of titanium in alloy steels. **J. AERND AND H. SCHWELLENBRACH. *Arch. Eisenhüttenw.* 4, 245-7 (1930)**—The sepn. of Ti from Fe⁺⁺, Cr, Mn, Co and Ni is accomplished by pptn. with BaCO₃, filtering and washing the ppt. with hot AcOH. The remaining alloy constituents are removed by fusion with Na₂CO₃ and NaNO₃. The residue, contg. Ti as the metatitanate, is dissolved in concd. HCl, the soln. made ammoniacal and traces of Fe and Cu removed by pptn. with H₂S in the presence of tartaric acid. Finally, the Ti is pptd. with o-hydroxyquinoline, ignited and weighed as TiO₂. Colorimetric estn. may be substituted for the gravimetric after sepn. of the alloy constituents. If only Cr and Ni are present the colorimetric method eliminates the sepn. with BaCO₃ as the former may be removed by decantation as chromate and Ni up to 5% and Cu up to 0.5% have no effect on the detn. **H. F. JOHNSTON**

Determination of chromium in special steels. **P. KLINGER. *Arch. Eisenhüttenw.* 4, 7-15 (1930), *Stahl u. Eisen* 50, 1166-7 (1930)**—The Cr contents of 12 special steels with 1.5-12% Cr have been detd. by 9 different methods. Methods involving alkali fusions give the best results when the Cr is detd. iodometrically, and of those based on soln. of the steel in acids followed by oxidation to chromic acid the persulfate-AgNO₃ procedure is the most satisfactory. The effect of other metals on the results given by the various procedures is briefly discussed. **B. C. A.**

Data on the determination of boric acid. **G. VASTAGH. *Magyar Chem. Folyoirat* 37, 55-71 (1931)**—The titration of H₂BO₃ in neutral solns. contg. 1% of mannitol was studied. Acids (except H₂PO₄ and SiO₂) do not interfere if neutralized previously with KOH in the presence of methyl red. It was found that 2 distns. with H₂SO₄ and MeOH served to remove H₂BO₃ from cations, 0.06 mg. of H₂BO₃ could be detected in 2 g. of

$\text{Al}_2(\text{SO}_4)_3$ crystals The macro and microchemical detn of H_2BO_3 is described in detail S S DE FINAY

A rapid and sensitive method for the volumetric estimation of phosphoric acid. R BIAZZO *Ann chim applicata* 21, 75-81(1931)—Phosphate is detd volumetrically by adding an excess of a standardized soln of $(\text{NH}_4)_2\text{MoO}_4$ and titrating the excess with $\text{Pb}(\text{OAc})_2$ soln These solns are standardized by means of solns of known phosphate concn Tannic acid is used as an outside indicator for the titration Amounts of phosphate from 0.1 to 1.0 cg may be detd within less than 2% by the above method A W CONTIERI

The measurement of cuprous oxide with permanganate solution G BRUNING *Centr Zuckerind* 39, 13(1931)—The possible advantages of Müller's soln in the detn of invert sugar are discussed J F LEFF

Manometric determination of hydrogen peroxide. AKIHI FUJITA AND TAKESHI KODAMA *Biochem Z* 232, 15-9(1931)— H_2O_2 is detd manometrically either by decompg it in a strongly acid medium with KMnO_4 or with catalase By using vessels of 15-20 cc capacity the O evolution X_0 is measured by Warburg's method from the rise in level of the manometer fluid By both methods an accuracy of about 3% is attained S MORGULIS

The determination of moisture in liquid sulfur dioxide. A L FLENNER AND W R CAVERLY *Refrigerating Eng* 21, 344-5(1931)—The moisture in SO_2 is detd by passing the SO_2 through P_2O_5 contained in a chain of Fleming absorption jars The method may be used for the detn of very small amts of SO_2 of the order 0.002% The exptl error is not greater than 0.0002% absolute A H JOHNSON

Use of buffered ammonia in the iodometric thiocyanate determination. H ARMIN PAGEL AND HERMAN J KOCH *J Am Chem Soc* 53, 1774-8(1931)—To the neutral soln of CNS⁻ in a vol of 150-300 cc, add NH_4OH until faintly alk to litmus if heavy metals forming insol hydroxides with NH_4OH are present As buffer soln add 1 g of NH_4 sulfate, chloride or nitrate dissolved in 20 cc of N NH_4OH and treat with standard I soln, using a min excess of 5 cc After 4-5 min add 5 cc of 0 N HCl and titrate with standard $\text{Na}_2\text{S}_2\text{O}_3$ Before acidification the products of the reaction between CNS⁻ and I_2 are SO_4^{--} , I^- , H_2O and ICN but upon acidifying the ICN , reacts with I^- liberating free I_2 so that the complete reaction can be expressed by the equation $\text{CNS}^- + 3\text{I}_2 + 4\text{H}_2\text{O} \rightarrow \text{SO}_4^{--} + \text{CN}^- + 8\text{H}^+ + 6\text{I}^-$ Certain modifications are necessary if KIO_4 is used as a source of I_2 W. T. H.

Determination of mercurous ions with potassium chloride solution. L NE ZOMBORY *Acta Sci Univ Francisco Iosephinae, Acta Chem, Mineral, Physica* 1, 1-4(1928)—A 0.4% soln of Na alizarinsulfonate is added to a 0.1 N soln of KCl and is used in titrating the Hg^+ soln to be tested The results agree with gravimetric data S S DE F.

Reactions of harmine. O FERNÁNDEZ AND P E RAURICH *Anales soc españ fis quim* 29, 74-6(1931)—Harmine is oxidized to m -nitroanisic acid, which is reduced to amine, then converted into a diazo compd which gives a purple color on treatment with β -naphthol Another test is the development of a carmine color when hydroharmine, obtained by hydrogenation of harmine with Na and EtOH , is acted upon by p -nitrodiazobenzene chloride E M SYMMES

Quantitative estimation of mixtures of isomeric unsaturated compounds. III. Review of the iodometric methods and a new bromometric method. REGINALD P LYNSTEAD AND JASON T WM MAYN *J Chem Soc* 1931, 723-5 In a mixt of 3 C tautomerides, the iodometric method is likely to require modification because the introduction of negative groups into the β - γ acids greatly diminishes the activity toward I For the analysis of a mixt of itaconic and mesaconic acids, consistent results can be obtained by the following empirical method provided the temp is kept const and the titrating reagent is standardized in the same way against known mixts of these acids. As reagent, 0.05 N Br dissolved in 40% KBr soln is recommended First det. the Br equiv by adding 25 cc. of the reagent to 10 cc. of water in a glass-stoppered bottle and after 10 min in the dark add 10 cc. of 10% KI soln and titrate with 0.05 N $\text{Na}_2\text{S}_2\text{O}_3$ In titrating known and unknown mixts of itaconic and mesaconic acids, use 10 cc. of 0.067 N soln of the acids in place of the H_2O and add an equiv quantity of Br . A temp of 16.8° was used in testing the procedure W. T. H.

Detection of very small quantities of acetylene. E PIETSCH AND A. KOTOWSKI *Z. angew. Chem* 44, 309-12(1931)—By means of a soln prepd by dissolving 1 g. of blue vitriol in 8 cc. of 10% NH_4 soln, adding slowly 3 g. of NH_4OH HCl and dilg to 50 cc., it is possible to detect as little as 3.7×10^{-10} by vol, or 1.7×10^{-6} g. of C_2H_2 in 4 l. of gas which is allowed to stream through the reagent under normal pressure at the rate of 1.7-1.9 l. per hr. The sensitiveness of the test is increased by adding 3-5 sq mm of

filter paper with smooth edges, the colloidal ppt. of $\text{Cu}_2\text{Cl}_2 \cdot \text{H}_2\text{O}$ being adsorbed by the filter paper and made more visible. W. T. H.

New gravimetric determination of benzidine and its homologs. Some new complex salts of these bases. W. HERZOG. *Chem. Ber.* 55, 117-120 (1922). A complex salt is produced by adding 7% H_2Cl_2 , H_2SO_4 or HCl soln. and filtering the ppt. on a sintered glass filter. Results with benzidine and o-toluidine were satisfactory. S. S. DE FINAIA.

The estimation of pyruvic acid. R. H. KERRICK. *Proc. Roy. Soc. (Edinb.)* 1925, 164. A method suitable for estn. of small quantities of pyruvic acid (2-20 mg.) in bud fluids consists in reducing the acid by Zn and H_2S in the presence of a trace of CuSO_4 , and oxidizing the resulting lactic acid to Ac^+ which is estd. by Clausen's method of titration against I^- . By means of the condensing unit used the aldehyde can be transferred into the receiver by reaction with CO_2 without any water passing into the receiver. I. J. C.

Reactions with Na nitrososulfate (TAVLINS, G. R. (NIDBACH, NIDBACH) & THIOXANATE ION) & Fe (NIDBACH, VAN VALKENBURGH & THIOXANATE ION) & Fe (NIDBACH, VAN VALKENBURGH & THIOXANATE ION). The separation of elements in vanadium's small quantities. TERRACHER, PHILIPP. 2. The use of the Sn^{2+} electrode in the determination of the concentration of H^+ ions and in potentiometric titrations (BARTON, ROBERT). 2. The solubility of La oxalate and of La hydroxide in water (KOLTHOFF, IANETIS). 2.

SCHMITZ, P. Gassanalyse in der Technik. Leipzig. S. Hirzel. 79 pp. M. 5.

Titration tables. FAY H. GUTENYER (to The Electric Smelting and Aluminum Co.) Can. 210,651, Apr. 21, 1931. The oxidation value of a perchlorite soln. is found by rendering the soln. acid in the presence of available combined I^- and measuring the quantity of free I^- released by the addition to the soln. of tablets, each contg. a standard quantity of a thiosulfate of an alkali metal with NaHSO_3 . The tablets are added one by one at intervals long enough to enable decolorization of the tablets and agitation of the soln. by the escape of gas from the soda, whereby the strength of the hypochlorite soln. will be indicated by the number of tablets required to eliminate the color caused by the free I^- .

8-MINERALOGICAL AND GEOLOGICAL CHEMISTRY

EDGAR T. WHERRY AND J. F. SCHAEFER

Sulfur crystals from Ajka and Pilsb-Szent-Istvan. SANDOR KOCH. *Ann. Hist. Nat. Mus. Nat.-hist. Hungaria* 25, 451-51 (Hung.), 434-5 (Ger.) (1928). *Mineralog. Abstr.* 4, 520. Small S crystals, showing 13 forms, were produced by burning heaps of lignite. J. F. SCHAEFER.

A notable example of Spanish pyrite. F. DIAZ TOSATS. *Mem. soc. espal. hist. nat.* 15, 13-5 (1920). *Mineralog. Abstr.* 4, 520. Acicular pyrite crystals occur at Mercedal. J. F. SCHAEFER.

Notes on some Spanish arsenopyrite crystals. JULIO GARRIDO AND CARLOS VALDEAVELLANO. *Mem. soc. espal. hist. nat.* 15, 422-54 (1920). *Mineralog. Abstr.* 4, 521. Crystallographic. J. F. SCHAEFER.

Some bismuth minerals of Vaskő (Hungary). SANDOR KOCH. *Matematik. es Természettud.-munk. Értesit.* 47, 210-23 (German abstract 227) (1930). Analyses of bismuthinite gave: S 19.45, 19.62, Cu 0.57, 0.93, Pb 0.69, 5.29, Fe 0.42, 0.11, Bi 80.04, 76.74 SnO_2 —, 0.45 sums 100.16, 100.12%. Cosalite showed: Ag 0.52, Cu 3.41, Pb 57.65, Fe 0.68, Bi 41.75, S 15.82, sum 99.71%. formula $\text{Pb}_2\text{Bi}_2\text{S}_4$ or $2\text{PbS} \cdot \text{Bi}_2\text{S}_3$ (Fe content deducted as chalcopirrite). Analyses of rechenite gave: Cu 4.17, 4.04, Pb 1.35, 15.10, Fe 0.46, 0.42, Bi 59.28, 59.22, S 17.53, 17.54, SnO_2 —, 0.12, sums 100.14, 99.91%, corresponding to $\text{CuPb}_2\text{Bi}_2\text{S}_5$ or $\text{CuS} \cdot 3\text{PbS} \cdot 5\text{Bi}_2\text{S}_3$. One-mm balls of bismuthosphenite were found on bismuthinite. S. S. DE FINAIA.

Ramdohrnte, a new mineral from Bolivia. FRIEDRICH AUFSELD. *Ann. Minera. Geol.* Abt. A, 1930, 365-7. *Mineralog. Abstr.* 4, 541-2. Ramdohrnte is found as bluish, grey-black prisms in quartz with accompanying pyrite, stannite, sphalerite, jamesonite and pyrrhotite at the Chocaya la Vieja mine, Potosi. They show "alternation twinning", sp. gr. = 4.15, hardness = 2. Analyses gave: S 2.11, Sn 31.1, Bi 2.5, Pb 30.3, Ag 10.1, Cu 0.5, Fe 0.3, gang. (SiO_2) 5.5, sum 100.2%, formula $\text{AgS} \cdot 3\text{PbS} \cdot 5\text{Bi}_2\text{S}_3$, near fireclite. J. F. SCHAEFER.

Proustite and xanthoconite from Nagyág. L. TOKOBY. *Centr. Mineral. Geol.* 1930A, 117-23, *Matematik és Természettudományi Értesítő* 46, 644-51 (in German 652-6).—In the locality of Nagyág, Hungary, quartz like material has been mined, consisting primarily of nagyágite (Au 7.61, Pb 54.50, Fe 0.93, Te 17.89, Sb 8.62, S 9.10 and quartz 2.12%), quartz crystals, dolomite rhombohedrons, covered by native As, a little pyrite and chalcopyrite, proustite and xanthoconite. The last two have not been previously discovered in Nagyág, and are found but seldom in other parts of Hungary. They are fully described here with diagrams of crystal structure. J. PINCHACK

The relations of etch-hills on an etched sphere of fluorite. E. ERNST. *Festschr. Victor Goldschmidt, Heidelberg* 1928, 89-97. *Mineralog. Abstracts* 4, 353.—Minute etch-hills are present on the rounded trioctahedral faces produced by the solvent action of HNO_3 on a sphere of fluorite. J. F. SCHAIKER

Zamboninite, a new mineral species. FRANCESCO STELLA STARRABBA. *Boll. soc. geol. ital.* 48, 259-63 (1930). *Mineralog. Abstracts* 4, 249.—The new mineral, zamboninite, occurs as mamillary masses with gypsum in an old fumarole channel in the scoria of 1069 on Mount Rocca, Etna. The masses have a grey, stony crust, but the interior is white, friable and radially fibrous with d 2.98-3.00, n 1.405-1.411, very weak birefringence with straight extinction and + elongation. Analysis gave Mg 24.26, Ca 17.94, Mn not detected, Al 0.41, Fe trace, Na 0.31, K 0.05, F 55.57, H_2O 1.63, sum 100.17%, formula $\text{CaF}_2 \cdot 2\text{MgF}_2$. It is probable that the mineral was not formed by direct sublimation, but by the action of F vapors on descending solns. contg. Mg and Ca. J. F. SCHAIKER

Two mineral notes on Moravia. JISKRA LONDÝNOVÁ. *Věstník Státního Geol. Ústavu Českosloven. Rep.* 3, 125-30 (1929). *Mineralog. Abstracts* 4, 379.—Occurrences of fluorsite, calcite and aragonite are noted. J. F. SCHAIKER

Some minerals of the Broken Hill Lode. G. SMITH. *Chem. Eng. Mining Rev.* 23, 179-82 (1931).—S has been unable to verify the presence of cerargyrite, though other Ag haloids, embolite and iodyrite have been found to be present. Neither was cerussite found. Crystals of iodyrite were found deposited on 1895 slabs of limonite. In quantity pyromorphite ranked as the 3rd and last of the great Pb-producing minerals of the Lode. Calamine was mined but as part of the gang, it was not treated with the object of obtaining Zn. Erythrite and smaltite were found in small quantity. Cobaltite was the most important Co mineral and was plentiful at various periods from 1890 to 1898. Some of these deposits were of considerable economic value because of their high Ag content. H. C. PARISH

Mode of formation of spiral groups of dachiardite. G. D'ACHIAI. *Atti second. Linee* (Cl. Sci. fis. mat. nat.) [6], 9, 182-7 (1929). *Mineralog. Abstracts* 4, 318-9.—Crystallographic. J. F. SCHAIKER

Stannierite (cobaltic hydronide), a new mineral. A. SCHOEF AND V. CUVÉLIER. *Bull. soc. belg. géol. pal. hydrol.* 39, 74-82 (1930). *Mineralog. Abstracts* 4, 347-8.—Stannierite occurs in some quantity in the upper oxidized portion of a mineral vein at Mundungu, Katanga, Belgian Congo. Analysis of material separated as far as possible from intermixed malachite, hematite, kaolin and quartz gave Co_2O_3 68.72, Fe_2O_3 9.45, Al_2O_3 6.87, CuO 2.68, SiO_2 2.67, H_2O 10.15, sum 100.54%, corresponding to $(\text{Co}, \text{Fe}, \text{Al})_2\text{H}_2\text{O}$. This compn. approaches that of transvaalite. The mineral is readily sol. in HCl with evolution of Cl_2 . It is black and compact with, in places, a shining mamillary surface. It has black streaks, sp. gr. 4.137, hardness 4½. It is opaque with a granular or banded concentric texture. In reflected polarized light it is birefringent and pleochroic. The mineral is regarded as the crystal equivalent of the colloidal heterogenite, and a preliminary x-ray examn. by W. F. de Jong suggests that it is isomorphous with goethite. J. F. SCHAIKER

Etch-figures on calcite from Roith. JULIUS HOLZNER. *Festschr. Victor Goldschmidt, Heidelberg* 1928, 163-9. *Mineralog. Abstracts* 4, 354.—Isolated calcite crystals from a clay seam in a limestone quarry at Roith, Lower Austria, are deeply corroded. The etch grooves along cleavage cracks, etch pits, etch-hills and flow-channels are described. J. F. SCHAIKER

The nature of lublinit and its solubility in distilled water. STANISLAW J. THUGUTT. *Arch. min. soc. sci. Varsovie* 5, 97-104 (in French 105-7) (1929). *Mineralog. Abstracts* 4, 334.—Lublinit from Mt. Pulawsk on the Vistula gave on analysis CaO 54.56, Fe_2O_3 + Al_2O_3 0.18, insol. 0.76, CO_2 44.37, loss at 160° 0.21, sum 100.03%. This material finely powdered and heated with water in an autoclave for 72 hours at 233° gave a colloidal suspension containing 0.0137 g. CaCO_3 per 100 cc. This on evapn. deposited rhombohedra of calcite. J. F. SCHAIKER

Barytocalcite and its structural relations to other materials. B. GOSSNER AND

F. MUSSGANG *Centr. Mineral. Geol., Abt. A*, 1930, 321-8. *Mineralog. Abstracts* 4, 364.—Barytocalcite (sp. gr. 3.684) has a unit cell of dimensions $a = 8.15$, $b = 5.22$, $c = 6.58$ A U ($a:b:c = 1.361:1.1:2.05$, $\beta = 81^\circ 52'$) containing 2 mols of $\text{CaBa}_2\text{C}_2\text{O}_6$. The space group is C_2^1 . These cell dimensions are compared with those of NaNO_3 , CaCO_3 (calcite), dolomite and baryte. J. F. SCHAIKER

Alstonite and milarite—a contribution to the study of complex crystals. BAL-THASAR GOSSNER AND FRANZ MUSSGANG *Centr. Mineral. Geol., Abt. A*, 1930, 220-38. *Mineralog. Abstracts* 4, 365.—X ray photographs of bromite (alstonite) (sp. gr. 3.707-3.715) gave a unit cell of dimensions $a = 8.77$, $b = 4.99$, $c = 6.11$ A U, ($a:b:c = 1.757:1.1:2.24$) containing 2 mols of $\text{CaBa}_2\text{C}_2\text{O}_6$. The x ray data suggest that the mineral is a double salt (CaCO_3 , BaCO_3) rather than an isomorphous mixt. of aragonite (CaCO_3) and witherite (BaCO_3). Milarite gives data conforming with hexagonal symmetry and the space group D_{3h}^1 . The unit cell of dimensions $a = 10.46$, $c = 13.90$ A U, ($c/a = 1.329$) contains 2 mols (Si_2O_5) $_{1/2}$ $\text{Al}_2\text{Ca}_2\text{KCl}$. The crystals show optical anomalies with a division into orthorhombic sectors. J. F. SCHAIKER

The Fedorov method in the determination of feldspars. VIKTOR ZSIVNY. *Feldspars Ausz. 58*, 93-108 (1929).—The method and some details by it are given.

Determination of the species of plagioclase by the theodolite microscope. S. K. CHATTERJEE *Proc. 15th Indian Sci. Cong.* 1928, 288. S. D. DE FINLEY

Notes on minerals from the Canaries. RAFAEL CANDEL VILA. *Mem. soc. esp. hist. nat.* 15, 263-8 (1929). *Mineralog. Abstracts* 4, 314.—Crystallographic notes on olivine and hornblende. E. J. C.

The formula of eucolite. FELIX NACHATSCHKEI *Centr. Mineral. Geol., Abt. A*, 1930, 300-4. *Mineralog. Abstracts* 4, 358.—From Cleve's 1890 analysis of eucolite (eudialyte) from Norway the formula is deduced as $\text{X}_2\text{Y}_2\text{Si}_6(\text{O},\text{OH})_{11}$, where $\text{X} = \text{Ca}$, Na , Yt , Ce and $\text{Y} = \text{Zr}$, Fe , Mn , Mg , Cb . Four such mols. are contained in the unit cell of Gossner and Mussgang (cf. following abstract). J. F. SCHAIKER

The structural and molecular unit of eudialite. B. GOSSNER AND F. MUSSGANG *Centr. Mineral. Geol., Abt. A*, 1930, 81-3.—In an attempt to det. the mol. structure of eudialite, the authors studied the uniformity of the structure and mol. wt. by measuring the lattice structure of the crystals and by estg. the mol. wt., assuming eudialite to have 6 possible chem. compns. The density, $s \times M$, was assumed (suspension method) to be about 2.929. The elementary substance was found to be rhombohedral in character, the crystals showing the form (111), (100), (110), (101), (211), (111). The d of this substance was estd. at 3.145. The mol. uniformity could not be ascertained on the basis of the accepted structural uniformity, all 6 possibilities presented contradictions to exp't. work, e. g., the formula $2\text{Na}_2\text{SiO}_3 \cdot \text{CaFeSi}_2\text{O}_7 \cdot \text{ZrSi}_2\text{O}_7$ showed the best relationship, but did not agree with the structural uniformity. J. F. SCHAIKER

Remarks on the work of B. Gossner and F. Mussgang. WILLIAM H. ZACHARIASEN *Centr. Mineral. Geol., Abt. A*, 1930, 315-7. *Mineralog. Abstracts* 4, 367-8.—X ray rotation and Laue photographs of a crystal of eudialyte from Nauyasak, Greenland, gave a rhombohedral cell of edge $a = 13.01$ A U and axial angle $60^\circ 44'$. The space group is D_{3h}^1 . Splinters from the crystal had d ranging from 2.826 to 2.888, and they show differences in optical character, being isotropic or weakly birefringent with blue-violet interference colors, the material is therefore presumably variable in chem. compn. Z from later analyses suggests a formula $\text{R}_2\text{Si}_4\text{N}_{11}$ (where $\text{X} = \text{O}, \text{OH}, \text{Cl}$) or $(\text{Na}, \text{Ca}, \text{Fe})_2\text{ZrSi}_4\text{O}_{11}(\text{OH}, \text{Cl})$, for which there would be 4.13 mols. in the unit cell. J. F. SCHAIKER

Sulfatic cancrinite from the Ilmen Mts. (South Ural). A. N. ZAVARITSKH *Mém. soc. russe mineral.* [2], 58, 201-7 (1929). *Mineralog. Abstracts* 4, 379.—Blue cancrinite occurs as nests and veinlets in nephelite crystals in pegmatite veins in the nephelite-syenite of the Ilmen Mts. Two analyses gave SiO_2 34.78, 35.29; Al_2O_3 28.77, 28.79; Fe_2O_3 0.10, 0.19; MgO 0.12, 0.10; CaO 1.27, 1.49; Na_2O 15.59, 15.65; K_2O 4.90, 4.15; H_2O 7.01, 7.62; CO_2 0.89, 1.01; SO_3 6.25, 5.76; sums 99.68, 100.05%. The SO_3 content is rather high. J. F. SCHAIKER

Eulytite from Dognafeska. SANDOR KOCH *Matematik és Terméstudományi Értesítő* 46, 640-2 (German abstract 643) (1929).—The forms {100}, {111}, {111} and {211} were observed on 1 mm crystals. Larger crystals were of a hexahedron, the smaller of a triakistetrahedron habit. S. D. DE FINLEY

Mode of occurrence and chemical composition of garnet from Nellore district, Madras. V. S. SWAMINATHAN *Proc. 15th Indian Sci. Cong.* 1928, 288.—Garnet is a common constituent mineral of both the schistose and the pegmatitic rocks and often

occurs as an accessory of the granite and gneiss outcrops exposed in the country. Cryst. as well as massive forms are found, the latter being generally heavy, granular and of a comparatively large size. Dodecahedra, trapezohedra and combinations of these 2 are the most common cryst. forms. The dominant colors are red, dark red and varying shades of brown. The sp. gr. varies from 3.52 to 4.15. R. J. C.

Chemical composition and physical constants of apicarsite from Popple near Jihlava. RABIN NOVÁČEK. *Časopis Mineralogického Muzea* 27, 4 pp. (preprint) (1930). *Mineralog. Abstracts* 4, 350. This garnet has a sp. gr. 4.212, n_x 1.5040, n_y 1.8131, n_z 1.8175 and gave on analysis: SiO_2 36.81, TiO_2 0.12, Al_2O_3 20.70, FeO 0.78, FeO 12.11, MnO 29.00, MgO 0.14, CaO 0.27, H_2O 0.02, sum 100.31%; corresponding to spessartite 18.21, almandine 30.56, pyrope 0.47 and grossularite 0.73%. J. I. S.

The general formula of vesuvianite and its relation to garnet. FRITZ MACHATZCHKE. *Centr. Mineral. Geol.*, Abt. A, 1930, 281-91. *Mineralog. Abstracts* 4, 357. Published analyses of vesuvianite are calc'd to at. ratios and a general formula deduced as $\text{Ca}_2\text{Y}_2\text{Si}_2(\text{O}_3\text{OH})_2$, where $\text{Y} = \text{Ca}$ or Na , $\text{Z} = \text{Mg}$, Fe^{2+} , Fe^{3+} or Al and $\text{X} = \text{Si}$ with some Al . The unit cell of vesuvianite has dimensions near to those of garnet, indicating a close relation. J. P. SCHAIER.

Vesuvianite from the ejecta of Sabatini. MARIA PIAZZA. *Periodico Mineral. (Rome)* 1, 8, 17 (1930). *Mineralog. Abstracts* 4, 330. A detailed description. J. I. S.

The general formula of melilite. FRITZ MACHATZCHKE. *Centr. Mineral. Geol.*, Abt. A, 1930, 278-81. *Mineralog. Abstracts* 4, 357. From the analyses tabulated by Berman (*C. A.* 24, 4181) a general formula for minerals of the melilite group is deduced as $\text{XYZ}(\text{O}_3\text{OH})$, where $\text{X} = \text{Ca}$ or Na , $\text{Y} = \text{Mg}$, Fe^{2+} , Fe^{3+} or Al and $\text{Z} = \text{Si}$ with some Al . J. P. SCHAIER.

Mineralogical notes IV. T. HOOGE SMITH. *Rec. Australian Museum* 17, 409-12 (1930). *Mineralog. Abstracts* 4, 345-6. Muscovite, a new mineral from Broken Hill, New South Wales, is jet black and compact, very brittle, friable, with a yellowish brown streak, hardness 3, sp. gr. 2.051. Thin sections are pale brown and optically isotropic. It fuses with difficulty to a black magnetic glass and is decomposed by acids with sign of granular SiO_2 . Analysis gave: SiO_2 52.35, Al_2O_3 0.41, Fe_2O_3 10.22, MnO 25.18, MgO 0.65, CaO 2.10, H_2O (105°) 23.01, H_2O (140°) 1.02, H_2O (±110°) 4.11, insol. (SiO_2) 0.70, sum 100.08%, leading to the formula $1.00\text{Al}_2\text{Si}_2\text{O}_{10}(\text{OH})_2 \cdot 2\text{H}_2\text{O}$ or $1.00\text{Al}_2\text{Si}_2\text{O}_{10}(\text{OH})_2$. J. P. SCHAIER.

Zeolites in extrusive rocks—a contribution to the problem of "Sonnenbrand." K. HORTER. *Z. prakt. Geol.* 38, 17-20 (1930). *Fortschritte Mineral. Kryst. Petr.* 14, 37 (1929). *Mineralog. Abstracts* 4, 374-5. Many basalts, apparently uniform when freshly broken, show numerous pale patches after a short exposure to the air (Sonnenbrand). This behavior is due to a hydrous, apparently glassy groundmass, which has been supposed to be zeolitic in nature. The pale patches give a more continuous dehydration curve than does the normal basalt. By centrifuging the finely powdered material of the light patches in Clerici soln. of sp. gr. 2.45 the groundmass was approximately isolated. The material gave a definite base-exchange with NaCl and CaCl_2 solns. It concludes that these basalts contain an alkali glass which is readily decomposed by moisture to form a zeolite mixt. J. I. S.

Zeolites from Chibino Mountains and Luvavrut in Russian Lapland. A. N. LAMBERTOV. *Trava muséum Acad. Sci. U. R. S. S.* 2, 91-100 (1927). *Mineralog. Abstracts* 4, 373-4. Zeolites found in the Khibinsky and Lovozersky tundras, Kola, include natrolite, analcite, chabazite, heulandite and mesolite. Analcite is always associated here with aegirite and may be either primary or secondary after sodalite or aegirite. The following equations are suggested: $2(\text{Na}_2\text{Al}_2\text{Si}_2\text{O}_8 \cdot 2\text{H}_2\text{O}) + 2\text{NaCl} = 2(\text{Na}_2\text{Al}_2\text{Si}_2\text{O}_8 \cdot 2\text{H}_2\text{O}) + 2\text{H}_2\text{O}$ (analcite) + 2NaCl or 2NaOH . Further white masses of hydrated alumina are found. From aegirite of the second generation: $\text{K}(\text{Na}_2\text{Fe}_2\text{Al}_2\text{Si}_2\text{O}_8) + \text{Na}_2\text{Al}_2\text{Si}_2\text{O}_8 \cdot 2\text{H}_2\text{O}$ (analcite) or $\text{Na}_2\text{Fe}_2\text{Si}_2\text{O}_8$ (aegirite) + $\text{Al}_2\text{O}_3 \cdot \text{H}_2\text{O}$ + $2\text{H}_2\text{O} = \text{Na}_2\text{Al}_2\text{Si}_2\text{O}_8 \cdot 2\text{H}_2\text{O}$ (analcite) + $\text{Fe}_2\text{O}_3 \cdot \text{H}_2\text{O}$. J. I. S.

Leucophyllite. MIKLOŠ VERNÉ AND A. RUMWALTER. *Min. berg u. hüttenmänn. Allg. u. Hochsch. Berg u. Forst. Sapon, Ungarn* 1930, 300-71. Leucophyllite from Sopron Mountain was analyzed: SiO_2 75.64, TiO_2 0.15, H_2O 0.11, Al_2O_3 14.21, Fe_2O_3 0.12, FeO 0.17, MnO 0.01, MgO 2.67, CaO 0.48, Na_2O 0.22, K_2O 0.50, H_2O 2.10, H_2O 0.06, sum 100.40%. O'Nann parameters, Niggli axes and halber optical and petrographic data are given. S. S. D. LINDAL.

New localities for zeolites in Moravia. BRUNO KUČERA AND BOHĀNA NOVÁČKÁ. *Časopis Moravského Zemského Muzea* 25, 211-27 (1927). *Mineralog. Abstracts* 4, 318. A complete list of old and new zeolite localities in Moravia and Czech Silesia is given. J. I. SCHAIER.

Analite and laumontite from Murlo (Tuscany). GIACINTA CUNEO *Ann. Mus. nat. hist. Genova* 52, 413-4 (1928). *Mineralog. Abstracts* 4, 376.—Clear isometric tetrahedra of analite gave SiO_2 53.69, Al_2O_3 22.91, Na_2O 14.27 (by difference), H_2O 9.13%. Laumontite SiO_2 50.90, Al_2O_3 21.51, CaO 11.76, H_2O 15.22, sum 99.09%. At 100° the laumontite lost 3.80% H_2O and at 133° 4.15%. J. F. SCHAIKER

Nontronite from two southern Russia graphite deposits. D. P. SKRIDTCHENKO. *Centr. Mineral. Gen., Abt. A*, 1929, 47-55. *Mineralog. Abstracts* 4, 327.—Seams of nontronite occur in pyroxene-graphite-gneiss in the Stary-Krym graphite mine near Mariupol. The mineral is decomposed by concd. HCl with sepa. of SiO_2 . Analyses on impure material are given. The material has been derived, perhaps, from the direct decompos. of Fe-bearing silicate or may have been formed by the action of solns. of Fe sulfate (from decompos. of pyrite) on secondary kaolin. J. F. SCHAIKER

The geological history of coal. GEORGE A. HICKLING. *Can. Mining J.* 52, 428-9 (1931).—A brief graphic description. W. H. BOTNOR

A new mineral of the nontronite-besideite group. I. JA. MICKY. *Centr. Mineral. Gen. Abt. A*, 1930, 293-303. *Mineralog. Abstracts* 4, 342.—A new mineral called *crusante* was found in a vein in liparite-pumice-tuff in the ravine of the Chubchur river, near Mt. Elbrus, northern Caucasus. It is chocolate-brown, sp. gr. 2.28, hardness 2, optically uniaxial, $n = \text{about } 1.56$. Analysis gave SiO_2 39.74, Al_2O_3 14.44, FeO 7.44, FeO 6.90, MgO 5.11, CaO 0.53, alkalis 2.80, H_2O (110°) 17.69, sum 94.67%, the deficiency being H_2O (above 110°). Over H_2SO_4 there is a loss of 14.89% H_2O . The formula is given as $1\text{RO} \cdot 0.8\text{R}_2\text{O}_3 \cdot 0.1\text{R}_2\text{O} \cdot 2.7\text{SiO}_2 \cdot 5.5\text{H}_2\text{O}$. This mineral differs from those of the nontronite-besideite group in the presence of FeO, MgO and alkalis, in its lower n and absence of pleochroism. J. F. SCHAIKER

Mesolite. A. CATTANEO. *Ann. soc. ital. sci. nat.* 65, 104-14 (1929); *Mineralog. Abstracts* 4, 319.—Mesolite from Iceland was studied by optical, gonometric, chem. and x-ray methods. Analysis gave SiO_2 44.91, Al_2O_3 20.30, CaO 11.15, Na_2O 4.10, H_2O 13.24, sum 99.71%. Crystallographic measurements and Lauegrams indicate monoclinic symmetry. The optic axial plane is perpendicular to the elongation, the sign is +, $2V = 70^\circ$, dispersion is strong, $B_x, c = 9/16$ and birefringence is low. Lauegrams of natrolite and scolecite indicated orthorhombic and monoclinic symmetry, resp. J. F. SCHAIKER

The minerals of the Krumná sphaerocerites. FRANTIŠEK SLAVÍK. *Společ. Průmyslové Společnost v Moravské Opatovici, Moravské Opatovici* 4, 90-3 (1929). *Mineralog. Abstracts* 4, 323.—The succession of minerals in the sphaerocerite septaria in the coal mines at Krumná, Czech Silesia, is the same as that in those of the Kládno coal basin in central Bohemia, i. e., quartz, ankerite, calcite, sulfide ores, nacreite. An analysis of nacreite is given. J. F. SCHAIKER

Phillipsite of the basalt in the region of Lake Balaton. BÉLA MAURITZ. *Mathematika és Természettudományi Értesítő* 45, 657-61 (German abstract 602) (1929).—A crystallographic description is given of phillipsite found in company of aragonite, chabazite and calcite. S. S. DE FENLEY

Analysis of monazite from Jun-an, Korea. KEIJIRO KIMURA AND SAKAE SHINODA. *J. Chem. Soc. Japan* 52, 47-54 (1931).—Hitherto there being only partial investigations and fragmental statements as to monazite from Jun-an, Korea, a complete analysis was performed, giving the following results: $d_{100}^{25} 5.106$, $n = 1.73$, CaO 0.53, FeO 0.09, Al_2O_3 0.28, Fe_2O_3 1.65, CeO_2 28.25, Nd_2O_3 etc., 27.87, Y_2O_3 etc., 2.47, ThO_2 9.49, SnO_2 0.15, SiO_2 1.85, UO_2 0.15, P_2O_5 20.07, Sb_2O_3 0.06, H_2O 0.79, CO_2 0.23, sum 99.93%. By a spectral analysis by arc light and x-ray the existence of La, Ce, Pr, Nd, Sm, Gd, Th, Dy, Ho, Er, Tm, Yb, Lu and Yt is confirmed, and Eu and Sc probably exist. In addn. Ba, Ti, Nb and Ta are detected by arc spectrum. From the Pb, U and Th contents, the age is calculated as 1.5×10^4 years (formula of A. Holmes and R. W. Lawson). T. IRIE

New data on lepidolite from Rožná, Moravia. JOSEF SEKANINA AND JOSEF VYSLOUŽEN. *Acta Soc. Sci. Nat. Moraviae (Brno)* 5, Pt. 2, 23-31 (1928). *Mineralog. Abstracts* 4, 379-80.—Analyses of pink lepidolite from Rožná gave SiO_2 47.63, Al_2O_3 28.69, FeO trace, FeO 0.33, MnO 0.75, MgO 0.15, CaO 0.35, Na_2O 1.03, K_2O 11.29, Li_2O 4.87, F 6.81, H_2O (+105°) 1.28, H_2O (-105°) 0.31%, the analysis agreeing with $\text{H}_4\text{Li}(\text{K}, \text{Na})_3\text{Al}_4\text{Si}_6\text{O}_{26}\text{F}_{16}$. J. F. SCHAIKER

Crystals of magnesium sulfate hexahydrate (hexahydrate). V. V. DOLIVO-DOBOVOLSKI. *Mém. soc. russe mineral.* [2], 53, 3-63 (1929); *Mineralog. Abstracts* 4, 378.—In the Sakai salt lakes, Crimea, after the sepa. of NaCl, spear-shaped crystals of $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$ (detd. by analysis), which differ from epsomite in form, separate. Thick, tabular, monoclinic crystals of $\text{MgSO}_4 \cdot 6\text{H}_2\text{O}$ also occur. Optical and crystallographic data on hexahydrate and on artificial $\text{MgSO}_4 \cdot 6\text{H}_2\text{O}$ are given. J. F. SCHAIKER

Blödite from Kalusz. ANTONI LASZKIEWICZ *Arch min soc sic Varsovie* 5, 79-94 (in French 95-6) (1929). *Mineralog Abstracts* 4, 334—Crystallographic Optical constants, for Na light, are $\alpha = 1.4526$, $\beta = 1.4855$, $\gamma = 1.4869$, $2V = 69^\circ 24'$, $B_x(\alpha) c = -42^\circ 11'$. J. F. SCHAIER

Recent minerals from the mines of Karvinná. VACLAV MAŠTALÍK *Sborník Přírodovědecké Společnosti v Moravské Ostravě, Moravská Ostrava* 4, 183-90 (1929). *Mineralog Abstracts* 4, 326—The coal mines have yielded calcite, gypsum, mirabilite, epsomite and mixtures of aluminum and iron sulfates. The descriptions given, with incomplete chem analyses and no optical data, are insufficient to det. the species exactly. J. F. SCHAIER

Wulfenite crystals from Almalyk. O. G. PADUROVA *Mém soc russe minéral* [2], 58, 109-15 (1929). *Mineralog Abstracts* 4, 378—Microchem tests showed the presence of Mo, Pb and Ba in crystals of wulfenite from Almalyk in the Tyuya-Muyun district, T'ergana. J. F. SCHAIER

The minerals of the uranium deposit of Katanga. A. SCHÖEP *Ann musée Congo Belge, Ser I Mineral, Terruere* 1, No 2, 43 pp (1930). *Mineralog Abstracts* 4, 313—The Kasolo radium mine is on Kasolo hill, 22 km S S W of Kambove. Two kilometers to the N E is a hill called Shinkolobwe (or Chinkolobwe) in which there is a Cu vein. These 2 mines are on the same mining block. The U occurs in veins 20-30 cm thick in bedded dolomites and slates. Twelve new species have been described from this locality: viz. becquerelite, curite, dewindite, dumontite, fourmarierite, ianthinite, kasolite, parsonsite, renardite, schoepite, sikhodowskite and soddyite. Other minerals present are uraninite, torbernite, quartz, talc, magnesite, heterogenite, linnaeite, Au, wulfenite and monazite. J. F. SCHAIER

Whewellite from the tertiary strata of the Maykop region (northern Caucasus). N. B. VASSOEVICH AND N. K. RAZUMOVSKII *Mém soc russe minéral* [2], 57, 275-300 (1928). *Mineralog Abstracts* 4, 377-8—Yellowish tabular crystals of whewellite ($\text{CaC}_2\text{O}_4 \cdot \frac{1}{2}\text{H}_2\text{O}$) were found in crevices in calcareous concretions from bituminous Eocene clay near Khodyzhenskaya in the Maykop district. Sp gr 2 22, hardness $3\frac{1}{2}$, $\alpha = 1.49$, $\beta = 1.55$, $\gamma = 1.65$, $2V = 82\frac{1}{2}^\circ$, optically +. Analysis gave CaO 35.80, CO_2 46.92, H_2O 10.60, CaCO_3 6.68%. The mineral is partly replaced by quartz and CaCO_3 . A summary of the literature on whewellite is given. J. F. SCHAIER

A meteorite from Oesede near Osnabrück. K. BUSZ *Vierteljahrsschrift Naturw. Ver. Osnabrück* 21, 4 pp (reprint) (1929). *Mineralog Abstracts* 4, 258-9—Thin sections of the meteorite show the ground mass to consist of broken fragments of bronzite and olivine, and chondrules of the same minerals with a granular or eccentric radial structure. The stone is a spherulitic chondrite. Analysis gave: Fe 18.17, Ni 1.90, FeS 5.77, SiO_2 35.57, Al_2O_3 2.71, FeO 7.99, MgO 23.24, CaO 2.96, Na_2O 1.31, K_2O 0.40, P trace, sum 100.03%, corresponding with nickel-iron 20.07, troilite 5.77, feldspar 22 and olivine 52%. J. F. SCHAIER

Meteorite from Ojuelos Altos. L. FERNÁNDEZ NAVARRO *Bol. soc. españ. hist. nat.* 29, 19-24 (1929). *Mineralog Abstracts* 4, 259—Thin sections of the meteorite show hypersthene, hmfeldspar, olivine, much magnetite with a border of hematite and some glass. The stone resembles a basalt tuff and is classed as a hypersthene oligochondrite. Sp gr 3 54-3 89. Bulk compn by analysis: SiO_2 40.02, TiO_2 nil, Al_2O_3 2.77, FeO 11.34, MgO 25.34, CaO 2.49, Na_2O 0.76, K_2O 0.23, P_2O_5 0.34, H_2O (105°) 0.09, FeS 6.76, Fe 8.17, Ni 1.12, Co 0.04%. J. F. SCHAIER

The meteorite of Olmedilla de Alarcón (Cuenca). L. FERNÁNDEZ NAVARRO *Mém soc. españ. hist. nat.* 15, 859-65 (1929). *Mineralog Abstracts* 4, 259—The meteorite is classed as an oligo-sidiferous hypersthene-chondrite with a marked brecciated texture and veined, approaching the eucrites, but perhaps representing a new type. There are chondrules of hypersthene, monoclinic pyroxene and basic feldspar, and magnetite bordered with hematite is abundant. Sp gr 3 712. Analysis gave the bulk compn.: SiO_2 36.92, TiO_2 nil, Cr_2O_3 0.38, Al_2O_3 2.52, FeO 9.39, MnO 0.26, MgO 23.59, CaO 2.16, Na_2O 0.87, K_2O 0.19, H_2O (105°) 0.09, P_2O_5 0.31, FeS 4.66, Fe 17.06, Ni 1.58, Co 0.02%. J. F. SCHAIER

Our mineral wealth [Paraguay]. GENARO ROMERO. *Rep. Paraguay, Cartilla Informativa* No 18, Asunción, 70 pp (1930). *Mineralog Abstracts* 4, 314—A brief sketch is given of the rocks and minerals of Paraguay with special mention and chem analyses of ores of Fe and Mn. J. F. SCHAIER

Notes on the iron ore of Mandi State. S. K. ROY *Proc. 15th Indian Sci. Cong.* 1928, 288—The unfossiliferous Krol slates of the Mandi state are characterized by the presence of a magnetite-chlorite-schist in which, at certain places, magnetite attains such an enrichment that the rock may then be considered as an Fe ore, magnetite-haematite-

quartz schist The transition between the ordinary Simla slates and the magnetiferous slate is gradual. Characteristic is the presence of a blue slate contg. beautiful cubes of pyrite just on both sides of the Fe-ore bed, which is about 400 ft. in thickness. The ore with the greatest concn. of Fe is found in the Chichot Subdivision of the state. Inferior types of magnetite-chlorite schist have been traced in the northern parts of the state near Barot, Pansaund, Dhamesth, etc. The better type of ore has been followed for about 12 miles between Pansar and Sangalwara, Chichot. Near Charan the strike direction has apparently been affected by the intrusion of the Mandi granite gneiss massif, which is therefore younger than the Krol beds. Magnetite, sepd. by washing from the disintegrated ore, on analysis showed 73.28% of metallic Fe with very little P and S, and a sample of Maroti ore gave about 61% Fe. This Fe deposit of Chichot was not reported previously. E. J. C.

Ilmenite and titaniferous iron ore from Nellore District, Madras. V. S. SWANATHAN. *Proc 15th Indian Sci Cong* 1928, 287.—Ilmenite occurs in the form of large segregated masses in a pegmatite vein, situated a mile N. E. of Gulmcherla village, Rapur Tq. The associated minerals are feldspar and biotite. Individual masses often weigh 30-50 lbs. The mineral breaks up easily into thin plates and tabular pieces, and the freshly broken surface reveals a submetallic luster. The sample gives, on analysis, 49.82% of TiO_2 . The mineral is available in fairly large quantities and could, with profit, be exploited, as it is extensively used in the manu. of ferrotitanium alloys and titanium white. Titaniferous Fe ore is met with in limited quantities closely assoc. with another pegmatitic vein near the Patragunta mica mine (Gudur Tq.). The specimens are tabular and have a dull steel gray luster. The results of the analysis of the mineral are: SiO_2 0.65, TiO_2 28.71, Al_2O_3 0.49, FeO 0.00, FeO 67.71, MnO 1.89, MgO 0.90, CaO 0, P_2O_5 0, sum 100.34%. Only very limited quantities of the mineral are available. E. J. C.

Climax molybdenum deposit of Colorado—with section on history, production, metallurgy and development. B. S. BUTLER, JOHN W. VANDERWILT AND CHAS. W. HENDERSON. *Proc Colo Sci Soc* 12, 311-53 (1931).—The broken ore reserve (Jan. 1931) is sufficient to furnish 2000 tons daily for 3 years, and plans are under way for increasing production with the demand. The Mosquito fault seps the Paleozoic rocks of the western part of the district from the pre-Cambrian rocks to the east, the Climax Mo deposit being in the granite of the latter. The mineralized area includes a central core, largely quartz, and an envelope consisting of moderately altered rocks cut in all directions by closely spaced veinlets. The molybdenite ore occurs in a zone of the envelope consisting of silicified rock surrounding the central core. Assoc. minerals are pyrite, a little chalcopyrite, brown sphalerite, hubnerite, topaz, quartz and fluorite. W. L. HILL.

Mineral resources of Palestine and Transjordan. G. S. BLAKE. *Jerusalem* (Pub. Works Dept.) 1930, 41 pp., *Mineralog. Abstracts* 4, 311.—A short outline of the geology is given. The most important mineral product is the Dead Sea brine, which is rich in KCl, MgCl, and MgBr₂. Exptl. production of salts by solar evapn. in large pans yielded first NaCl, then carnallite and last MgCl₂ with bromide (MgBr₂ 1.46%). Pure KCl can be separated by recrystn. from the carnallite. Petroleum, bitumen and bituminous limestone are of some importance in the region. In Transjordan occur small deposits of Cu and Fe ores and considerable gypsum and phosphate. J. F. SCHAIERER.

A deposit of barite in Orchha State, Central India. M. K. ROY. *Proc 15th Indian Sci Cong* 1928, 289.—In the Orchha State in Central India, in the village Suraypur, formerly called Khura, occurs a deposit of barite in Bundelkhand gneiss. The vein is traceable for more than 0.25 mile in length and is over 8 ft. in thickness. It is a quartz barite vein, in which barite appears to have been deposited later, filling up the cavities of the quartz vein and, in places, also of the adjoining country rock. A little Cu pyrite and malachite are found in the vein but no Pb ore. The deposit is of hydrothermal origin and apparently belongs to a series of several pyritic ore bodies occurring in the neighborhood and all having the same N. E.—S. W. direction. E. J. C.

The origin of coal. WALTER FUCHS AND OTTO HORN. *Z. anorg. Chem.* 44, 180-4 (1931).—The opinions of different investigators on the origin of coal are discussed, with a repetition of some exptl. work. Some new data are given, and numerous tables for comparison are added. A. HIRSCH.

Geologic survey of the environments of the radium mine of Tyuya-Muyun. D. I. MUSHKETOV. *Trav. radium et minerais radioactifs acad. sci. U. R. S. S. II*, 3-13 (1928).—The only dependable and scientifically recognized deposit of radioactive minerals in the Fergana Co., formerly Skobelev Co., is the deposit of ores contg. V, U, Ca and Cu in the region of Tyuya-Muyun. The residue from this ore after treatment contains a con-

siderable amt of Ra salts. Several hot springs which are located along Kisi Kungei-Karatash have a temp of about 16° with an emanation content from 1 to 2.21 Maché units. Hot springs of the southern slope of Tyuya-Muyun have temps from 21° to 25° with an emanation content from 0.1 to 0.3 and finally other hot springs with a temp. of 18.2° and 20.2° contain the max amt of emanation, 6.26 and 4.17 Maché units, resp. The temp of all of the hot springs exceeds 16° while the av yearly atm temp could not exceed 10° .

A C No6

The paleozoic deposits in the valley of the Aravan River. D V NALIVKIN *Trav. radium et minerais radioactifs acad sci U R S S II*, 14-20(1926) —The Aravan River crosses the Tyuya-Muyun ridge which contains deposits of radioactive minerals on the walls of many of its caves. The mass of the ridge consists of Silurian, Devonian and Carboniferous rocks. The specific radioactive minerals which are found in the walls of the caves are discussed in another paper.

A C No6

Petrographical particularities of Tyuya-Muyun radioactive deposits in Fergana. V I LUCHITSKIY *Trav. radium et minerais radioactifs, acad sci U R S S II*, 21-50(1926), cf *C A* 17, 3659 —At the beginning of the Kurgis Atan River are strongly developed sediments which are of light gray color. They are formed by an aggregate of quartz and feldspar of uniform grain structure. Quartz is found in small irregular grains. Microcline is represented by large irregular grains which are evenly distributed throughout the strata. Plagioclase and oligoclase appear in the form of sharply defined plates with (010) cryst grains and are included in microcline. Brown biotite, which often is transformed into chlorite or green hornblende, occurs in large quantities. These deposits are of Lower Carboniferous age.

A C No6

The eruptive rocks of the Tyuya-Muyun district. I. F AINBERG *Trav. radium et minerais radioactifs acad sci U R S S II*, 51-72(1926) —The Tyuya-Muyun Ridge is located in southeast Fergana in the region of the northern approach to the Alai Mountains. It is a part of an elevated plateau which extends to the South in the deep Naukat Valley. The entire district from Tyuya-Muyun up to the northern edge of the plateau is composed of Paleozoic slates, which are either bright slints or jasper, or of black carbon-like slates. These deposits are characteristic deep sea deposits. In some places one could find bituminous limestones, layers of black slint slates and layer like deposits of diabasic porphyrites. The following primary and secondary mineral components were found by microscopic examn: (1) rock-forming minerals Augite and feldspar, (2) accessory minerals, contained in the compn of the main strata Titanic iron oxide, pyrite, magnetite, apatite; (3) secondary minerals calcite, chlorite and quartz, most commonly found, sericite and epidote, found but seldom. A very peculiar characteristic is common to all those types where feldspar is found. The feldspar appears in all cases as albite.

A C No6

The gypsum in the calamine deposits of Sardinia. PIERO BONATE *Riscontri assoc mineraria sarda* 36, No 1, 39-41(1931), cf *Chimiglia, C A* 25, 809 —Gypsum is rarely present in the deposits of Sardinia, because it was formed in earlier periods and then carried away by the action of the waters. This action is also shown by the existence, at the lowest levels, of sedimentary rocks, very similar to the alluvial deposits. The gypsum, if protected from the action of the waters, is found in well defined crystals accompanied by blende and calamine. Its formation was probably the following: the pyrite was oxidized to FeSO_4 and $\text{Fe}_2(\text{SO}_4)_3$, which reacted with ZnS to form ZnSO_4 . This, by reaction with CaCO_3 , gave ZnCO_3 and CaSO_4 .

G. A. BRAVO

Materials related to the distribution of radioactivity in the western part of Fergana. A. N. SNEZAREV *Trav. radium et minerais radioactifs acad sci U. R. S. S. II*, 121-200(1926) —Every deposit of radioactive ore changes the ionization of air or its electric cond. Therefore, a study was made of these 2 valuable sources of information to locate radioactive ores.

A C No6

Investigations in the region of limestone contacts of the Kirghiz-Ata ridge in Fergana. V I LUCHITSKIY *Trav. radium et minerais radioactifs acad. sci U. R. S. S. II*, 201-63(1926) —The purpose was to det the geological and mineralogical character of the limestone contacts. These contacts are of interest, because certain traces of Cu ore are found, and the Cu ore in turn is usually found in connection with U compds. Particular minerals found are hornblende, porphyrite and hornblende-granites. Along the line of contact of syenitic-dioritic magma which formed the ridge of Kirghiz-Ata with Paleozoic limestones pronounced changes took place, as well as within the limestones which were exposed to different changes, as well as in the magma which in the middle part of the ridge gave birth to syenites and at the ends of the ridge gave birth to various diorites, among which the most important roles are played by hornblende and diorites. The first minerals formed during this process were ferric calcite rocks. Later, there

appeared plagioclases. Considerably younger are the ore minerals as, for instance, magnetite. During contact metamorphism no change took place which was due to the influence of pressure. The chief role has been played by the high temp. of the magma. Secondary changes in contact minerals took place chiefly in granite, which gave during its transformation 2 new secondary minerals: serpentine and chlorite. During these changes, hydrothermic agents were of considerable importance. A. C. NOB.

Notes on rocks from Sihor Hills, Bhavnagar, Kathiawar. V. S. SWAMINATHAN. *Proc. 15th Indian Sci. Cong.* 1928, 291.—Thirty hand specimens collected from the Sulidhar, Lamdhar and Tharungra hills situated west of Sihor town were examined. Bar ring a few dike rocks, rocks of a more or less acidic composition (differentiation products from a basaltic magma) seem to be prevalent in the area. The more interesting types are rhyolites, pitchstones, etc. E. J. C.

Some rocks of the Rajmahal Hills. P. C. DATTA. *Proc. 15th Indian Sci. Cong.* 1928, 291.—A study of rocks collected from a group of small conical hills called Gandsari in the Rajmahal Hills. These rocks were first described by Ball as trachyte and subsequently as andesite by McMahon. From a phys. and chem. study of the specimens D. has reached the conclusion that the rock is a basalt. E. J. C.

The volcanic rocks of the Irrawaddy Delta, Mangmya District, Lower Burma. H. L. CHUBBER AND M. M. WADSWORTH. *Proc. 15th Indian Sci. Cong.* 1928, 290.—The area which has been geologically mapped is situated on both sides of the Bassein river. On the east of the river the volcanic rocks form a plateau with an average height of about 100 feet presenting a steep scarp toward the river. The geology of the area has been described as follows: (5) Alluvium. (4) White and greenish fine tufts. These tufts enclose purplish or grayish nodules and boulders of trachytic lava (olivine-trachyte). It has been suggested that the eruptions were of a very explosive nature. (3) Coarse-grained grayish tufts with *Nummulites*. (2) Grayish volcanic tufts and ashes. These rocks were seen exclusively developed in the hills situated east of the village of Zyat Chaung. (1) Nummulitic limestones, sandstones and shales. The limestones are remarkable for containing a number of *foraminifera*. The area represents the southernmost occurrence of volcanic rocks in Burma situated on the igneous line connected with the Arakan Yomas and their continuation to the north and the south where numerous patches of serpentine occur. E. J. C.

The hornblende lamprophyres and associated rocks of Mokpalin quarries, Thabon district, Lower Burma. H. L. CHUBBER. *Proc. 15th Indian Sci. Cong.* 1928, 290.—The petrography of the rocks is as follows: (a) Contact rocks, e.g., veins of epidote rock containing a little quartz and feldspar. (b) Quartz veins, occurring as irregular veins or lenses. (c) Acidic pegmatites and apatite veins. (d) Lamprophyres. The common rock is a hornblende lamprophyre (camptonite) merging sometimes into a dark quartz-diorite-gneiss. (e) Hornblende-biotite gneiss, quartz diorite-gneiss and biotite-granite. E. J. C.

Magma types in the Deccan Trap. K. K. MATHEW AND V. S. DUBEY. *Proc. 15th Indian Sci. Cong.* 1928, 291.—The authors have studied a number of occurrences of igneous rocks, usually regarded as part of the Deccan Trap, in Gujarat, Kathiawar and Cutch. These can be classified into 3 groups: (1) the mafic or basalt magma type. (2) the mafic or ultra basic magma type and (3) the felsic or acid magma type. The first group is represented by the common Deccan Trap of the typical plateau basalt facies ranging in the later stages of eruption to an alkali rich basalt. This was followed by intrusions of highly basic rocks. The acid type appeared much later and is represented by rhyolites, granophyres, etc., in various parts of Western India. These groups appear not to have any immediate genetic relationship through a process of magmatic differentiation, but probably have their source deep down in distinct zones in the earth's crust. They have been traced in South India, the Himalayas, Burma and Western Asia. E. J. C.

Petrochemical data of the environment of Szarvaskő (Hungary). ZSIGMOND SZENTPÉTERI AND KÁLMÁN EMLÉNY. *Földtani Közlemények* 58, 109-14 (in German 216-22) (1928) (Published 1929).—Data for 12 rock analyses are given with a short explanation of conclusions drawn from them. S. S. DE FINÁLY.

Recent lava types of Etna. E. LENGVEL. *Földtani Közlemények* 59, 26-34 (1929).—Lava of the 1928 eruption of Etna is an extreme in lava types, being the most basic rock of all eruptions. It consists of an olivine-labradorite basalt with significant content of nepheline. S. S. DE FINÁLY.

Granites in the region of Mórógy. F. PAPP AND R. REICHERT. *Földtani Közlemények* 59, 35-41 (1929).—A petrographical study combined with 3 analyses of rocks from Mórógy (Hungary). S. S. DE FINÁLY.

Oligoclase rocks in the region of Szarvaskő (Hungary). ZSIGMOND SZENTPÉTERY. *Matematik és Természettudományi Értesítő* 47, 432-65 (in German 466-7) (1930) — Rocks consist of 70-80% oligoclase feldspars. Seven analyses are published with strikingly large Na_2O values, 8.05 to 9.12% (K_2O content being 0.30 to 0.75%). Rocks show many similarities to the gabbroidal mass of the environment. Analogous rock occurrences are found in Transylvania, at Koswa in North-Ural, in Central Asia, at Mariposa, Coalings, Colorado, etc. S. S. DE FINÁLY

Petrographic examination of the paleolithic splittings of Cave Budapest, Borsod (Hungary). AALDAR VÉNDL. *Matematik és Természettudományi Értesítő* 47, 468-83 (German abstract 484) (1930) — Splittings consist mostly of chalcedony, contg generally fibrous chalcedony. The occurrence of crystals of calcite is characteristic. S. S. DE FINÁLY

A study of the eruptive rocks of Lammersdorf (Hautes Fagnes). PAUL RONCHESNB. *Ann. soc. sci. Bruxelles Ser. B* 51, 52-7 (1931) — The rock of Lammersdorf is apparently similar to the rock of Greece. They differ especially by larger cryst. texture, less abundance of biotite and the presence of large crystals of plagioclase in recurrent zones. In general, the rock of Lammersdorf is the equiv. of the rock of Greece. Attention is called to the identical situation of the 2 eruptive masses of Greece and of Lammersdorf, on the north side of the anticline. These rocks are evidence of a contemporaneous intrusion and elevation by the same geodynamic process. This study is especially directed to chem. and microscopical examn. of the rock. Tables of chem. compns. are given, and mineral constituents are classified. ALICE W. EPPERSON

Andesite bomb with bread-like crust from Tusnádfürdő. VIKTOR ZSIVNY. *Matematik és Természettudományi Értesítő* 46, 277-90 (in German 291-3) (1929) — Microscopical and crystallographical examn. of a volcanic bomb of Tusnádfürdő (Transylvania). Chem. compn. (analyzed by K. Lmszt) SiO_2 65.15, TiO_2 0.44, CaO 4.41, MgO 2.19, FeO 2.85, Fe_2O_3 1.57, Al_2O_3 16.13, Na_2O 3.02, K_2O 2.14, SrO 0.08, H_2O (-110°) 0.19, and H_2O ($+110^\circ$) 1.33, total 99.51%. S. S. DE FINÁLY

Origin of the mica-pegmatites of Kodarma (Hazaribagh). S. L. BISWAS. *Proc. 15th Indian Sci. Cong.* 1928, 289 — The country rock of Kodarma constitutes the mesozone of the metamorphic belts and consists of mica-schist with subordinate amts. of sillimanite, andalusite, etc. The proportion of the last minerals increases with the depth as the katazone is approached. B. suggests that the residual plutonic magma which forms the pegmatites and was rich in volatile magmatic substances and rare earth compds. was initially at such a high temp. and under such a pressure as to be able to assimilate rocks of the deeper region, whereby it became specially rich in compds. producing muscovite, feldspars, etc. Injections through rocks other than mica-schist with sillimanite, andalusite, etc., do not appear to have such enrichment in the Kodarma mica area. This accounts for the occurrence of marketable mica in pegmatite veins within mica-schist and not in those cutting through other rocks. When the melt was forced up to its present position in the upper region, its thermal activity was so exhausted that no pyrometamorphism was effected. The magmatic emanation, however, produced marked pneumatolytic effect by the formation of minerals like tourmaline in the country rock. E. J. C.

Volcanic-ash showers. L. I. GRANGE. *New Zealand J. Sci. Tech.* 12, 228-40 (1931) — Showers of volcanic ash, each usually not more than a few feet thick, ejected from craters in the Rotomanga-Taupo-Tongariro zone, cover the central part of the north island of New Zealand. A general reconnaissance is given of the characteristics of the soil-forming showers. Tables show mech. and chem. analyses of air-dried soils from these areas. W. H. BOYNTON

The so-called ophiolitic rock of the mountains of Livornese and Castellina. G. DE GUIDI. *Boll. ist. super. agrario Pisa* 6, 53-68 (1930) — A geological, micro- and macroscopical description of this arenaceous rock and of its constituents. G. A. B.

Two new occurrences of crystalline limestone from Madura and Tinnevely districts. C. K. KRISHNASWAMY AND V. S. SWAMINATHAN. *Proc. 15th Indian Sci. Cong.* 1928, 298 — The authors think that the rocks were originally sedimentary in origin and have been subsequently subjected to metamorphism. E. J. C.

The siliceous oolites from the Cuddapah formation. L. RAMA RAO. *Proc. 15th Indian Sci. Cong.* 1928, 292 — A microscopic study of the oolites was undertaken with the special object of finding out whether any definitely recognizable org. remains could be detected as forming the nuclei of these oolitic grains. Several interesting microstructural peculiarities are shown by these oolites. One or two types of structures are very persistent, and the question is raised whether they may not be due to some very primitive organism. E. J. C.

A short note on the sedimentary petrography of some sandstones of the Rangaj stage from the Rangaj coal field. A. K. BANERJEE. *Proc 15th Indian Sci. Cong.* 1928, 291-2.—Four specimens of sandstone were examined microscopically and macroscopically. The result of the microscopic study shows the presence of quartz, feldspar, garnet, calcite, biotite and limonite. Portions of the hand specimens were crushed and sepd. by CHCl₃ and the minerals with a sp. gr. greater than 2.8 were, in the order of their frequency, garnet, zircon, rutile, barite, tourmaline, pyroxene, apatite and an unidentified Na-rich mineral. The 4 specimens of sandstones examined were obtained from the upper, middle and lower divisions of the Rangaj stage, and it is of the opinion that the heavy minerals do not furnish any data for distinguishing the beds from one another. Conclusion: The beds of the Rangaj stage were derived from the degradation products of a gneissic area with a granite intrusion, and there was no change in the direction of drainage during the Rangaj stage. E. J. C.

Petrography and coal classification (SEYLER) 21. Coal properties and its applications (KIRCKHOFF, SHIRANE) 21. Identifying processes from optical analysis of emerald rays (U. S. pat. 1,771,641) 1.

HATCH, F. H. Mineralogy. 6th ed., revised. London: Sir Isaac Pitman and Sons, Ltd. 292 pp. 6s. net.

MAIER, R., and JUNG, H. Die physikalische Chemie in ihrer Anwendung und Probleme der Mineralogie, Petrographie und Geologie. Revised ed. Jens G. Fischer. M. 10. Reviewed in *Z. physik. Chem.* 151A, 477-8 1930, *Chem. Abstracts* 10, 583 (1931).

TUCKER, FREDERICK G. The Examination of Fragmental Rocks. Stanford Univ., Calif.: Stanford Univ. Press. 157 pp.

9—METALLURGY AND METALLOGRAPHY

R. J. FENKREST, E. W. CHAFFET AND RICHARD KRECH

Secondary metals in 1929. J. P. DUNLOP. *Bur. Mines, Mineral Resources of the U. S.* 1929, Pt. 1, 174-201 (preprint No. 9, published February 24, 1931). E. H.

Lead in 1929. ELMER W. FLETCHER. *Bur. Mines, Mineral Resources of the U. S.* 1929, Pt. 1, 231-64 (preprint No. 11, published March 7, 1931). E. H.

Gold, silver, copper and lead in South Dakota and Wyoming in 1929. CHAS. W. HENDERSON. *Bur. Mines, Mineral Resources of the U. S.* 1929, Pt. 1, 205-73 (preprint No. 12, published April 29, 1931). E. H.

Iron ore beneficiation. CLYDE E. WILLIAMS. *Mining Met.* 12, 194-5 (1931).—Methods are surveyed and the possibilities of application to Lake Superior ores discussed. In 1930 over 15% of the ore of this district was beneficiated by washing, drying and screening and over 25% by crushing and screening. Beneficiation is increasing and should be further increased to conserve rich ores. Gaseous reduction to sponge iron is not recommended. The district will be a leader in iron ore production long after the date of predicted exhaustion. A. BUTTS.

Mineralogical aid in the concentration of manganese iron ores. W. H. COO-BELL, WARREN HOWES AND S. B. COCKE. *Eng. Mining J.* 131, 371-5 (1931).—Two manganese ores, representative of large deposits in the Cerritos district, in Minnesota, were studied by table concn., flotation, roasting and magnetic sepn., and mineralogical methods. The ores were quite similar in composition and appearance, but differed widely in results of sepn. by the methods mentioned. A microscopic study led to the conclusion that ore A did not give good flotation concentrates because of the presence of carbonates intimately associated with siliceous gang. Ore A responded to table concn. because the greater portion of the Mn and Fe minerals was fairly well liberated from the gang at table size. Ore B could not be tabled satisfactorily because of the intimate locking of the ore minerals with the quartz gang. Magnetic sepn. was impossible, because the Mn and Fe minerals were closely interlocked. Many photomicrographs and camera-lucida drawings are shown. W. H. BOYNTON.

Fine chemicals as precious-metal precipitants. E. T. ELLIS. *Metal Ind.* (London) 38, 377-8 (1931), cf. *C. A.* 24, 5004.—Chemicals that may be employed to ppt. Au are: NH₄ salts such as oxalate and sulfide, SbCl₅, FeSO₄, MgO, K salts, Na oxalate, SeCl₄ and ZnO. Other pptg. agents for Au and some of the other precious metals are formic acid, gallic acid, HgNO₃, oxalic acid, AcOEt, K citrate, KSCN, K tartrate, Na₂S, Na

tartrote, $\text{Na}_2\text{S}_2\text{O}_4$, tannic acid and tartaric acid. In all cases the chemicals should be of high purity. W. H. BOYNTON

Milling test of an oxidized gold ore from vicinity of Rockford, South Dakota. FLOYD C LLLINGSON. *Black Hills Eng* 19, 121-9(1931)—Amalgamation tests showed 50% Au recoverable by amalgamation in bottles, with prospects of larger yields by using a miller. Cyaniding on 100-mesh ore gave 96% recovery. The ore does not produce enough slime to interfere with percolation. Cyanide consumption was about 0.6 lbs per ton of ore, lime consumption was 20 lbs. H. C. DUSS

Group research among gold producers. R. M. P. HAMILTON. *Eng Mining J.* 131, 356-8(1931)—A plea is made for coöperation by producers in Northern Ontario. Direct operating costs and metallurgical results of 9 mills are tabulated and a flow sheet is shown for a 500-ton combination flotation and cyanide plant. Some of the successes of the group research plan in the Transvaal in Western Australia are touched upon. W. H. BOYNTON

The preparation and laboratory control of foundry sands. J. D. BURLIE. *J. Western Soc Eng* 36, 80-9(1931)—An outline is given of the general characteristics, source, spec., requirements and methods of control of foundry sands. Sands have 2 main classifications: those contg little or no natural bonding material and those contg it. The first group is used essentially in the production of synthetic molding sands and cores and the latter as a molding sand. Sand testing and control equipment are discussed and many data tabulated. Conclusions drawn are: The introduction of controlled properties of molding sand has contributed to the realization of a more positive control of casting technique by the use of suitable sands, the scrap losses may be reduced, and improved quality of castings results. Accurately controlled sand properties bear a close relation to an increased rate of production. W. H. BOYNTON

Natural gas for metallurgical furnace. BANCROFT GORE. *Black Hills Eng* 19, 103-12(1931)—The assay furnaces in the South Dakota School of Mines were converted from coal firing to natural gas firing. The advantages were found to be: greater speed of bringing cold furnace to working temperature, more even heating, control to within 10°F over any temperature between 500°F and 2500°F , low cost of operation (500 cu ft per hr as against 60 lbs coal), and possibility of checking up slovenly student work by metering gas used in furnaces. H. C. DUSS

The mathematical basis of the foundry shaft furnace, the determination of its size and mathematical regularity. ALBERT ACHENBACH. *Die Giesserei* 18, 217-25, 241-5(1931)—By making use of recent literature, math. relations involving the 4 principal factors in foundry shaft furnace practice (coke and Fe charged, amt, and velocity of blast, time of passage of charge and the height and capacity of shaft) are derived from considerations of theory and practice. It is shown that none of these factors can be changed arbitrarily without causing a corresponding change in the fusion process. J. BALOZIAN

Desulfurization [of pig iron] in the basic open-hearth furnace. K. KÖHLER. *Stahl u. Eisen* 50, 1257-64(1930)—The effect of CaO and Mn in the slag on the desulfurization of Fe in the basic open-hearth furnace has been studied. The S content of the finished steel is inversely proportional and that of the slag directly proportional to the CaO content of the slag. The rate of removal of the S is also a linear function of the CaO content of the slag between 25 and 40% CaO . A further increase in CaO increases the viscosity of the slag and, therefore, reduces its reactivity. No definite relation between the Mn content of the slag and its efficiency as a scavenger for S could be established. Slags with a low CaO and a high Mn content have no desulfurizing action, and a deficiency of Mn in no way impairs the desulfurizing action of the CaO . Addn. of a large excess of Mn does, however, appear to have some accelerating action on the removal of S from the metal bath. B. C. A.

The fundamental study for the production of low-carbon semi-steels from the standpoint of cupola operations. II. MASAYASU HORIKIRI. *Tetsu to Hagané (J. Iron Steel Inst Japan)* 16, 935-60, 1063-86(1930), *Metals & Alloys* 2, Abstracts 74; cf. C. A. 24, 45, 1605—II reports the result of an investigation on the production of low-C semi-steels on a large scale. Pig Fe and mild steel scrap were melted together in a large cupola. The temp. in the inner part was measured. The combustion gas at that place was sampled, and CO_2 , CO and O_2 were detd. The results thus obtained clearly showed that the C in semi-steels is absorbed by contact of red-hot coke with the molten Fe. It is, therefore, necessary to use the min. quantity of charged coke and to reduce the oxides in the charge with Si in order to obtain the low-C semi-steels. Moreover, the molten Fe contg. high Si dissolves only a small amt. of C, so that high-grade low-C semi-steels, generally, contain much Si. The oxidation of Fe by gases

gradually decreases from the tuyère level to the upper part, and increases on approaching the furnace wall but decreases at the center part. There is no significant relation between the depth of the cupola hearth and the C absorption of molten Fe. When many tuyères with small sectional area are used and the blast is passed through at high speed, the conditions are satisfactory, the atm. in the melting zone being oxidizing and considerable C being absorbed by the melt. On the other hand, with large sectional area and small nos. of tuyères the atm. is reducing, and the absorption of C decreases. The C content in the product depends on the contact condition of the gases with the melt and red hot coke at the melting zone; that is, the oxidation takes place by the gases, according to the following: $2\text{Fe} + \text{O}_2 \rightarrow 2\text{FeO}$, or $\text{Fe} + \text{CO}_2 \rightarrow \text{FeO} + \text{CO}$, and the reduction by the red hot coke, thus $\text{FeO} + \text{C} \rightarrow \text{Fe} + \text{CO}$, $2\text{FeO} + \text{Si} \rightarrow 2\text{Fe} + \text{SiO}_2$ and $\text{SiO}_2 + \text{C} \rightarrow \text{Si} + \text{CO}_2$. G. G.

The fusion process in the cupola furnace according to views and conceptions of technical investigators. CARL REIN. *Die Giesstete* 18, 33-7, 58-64, 78-83(1931)—A lecture. J. BALOZIAN.

Cupola malleable cast iron. H. H. SHIFFERD. *Foundry Trade J.* 44, 83-4, 102-4, 109, 125-9(1931)—A general discussion of the manu. of white heart and black heart malleable irons, their compn. and structure and also of molding sands used and methods of testing them for bond strength, permeability, etc. H. C. PARISH.

What reasons compelled the Prague Ironworks Company to introduce thin-walled blast furnaces? J. SAPP. *J. Iron and Steel Inst.*, Sept., 1930, Advance copy, 33 pp.—The theory of blast furnace working is discussed. In connection with the irregular working and severe corrosion of the walls it was observed that places where the lining had been melted off to only a few cm. thickness would last well for a comparatively long time. Five furnaces of the thin walled type have been erected on practically the original lines and have been eminently successful, as the lining has a long life, troubles in operation have been decreased, and the consumption of coke has been reduced, contrary to general conceptions. The original thickness of the lining of the latest furnace was 250 mm., and after 18 months the lining was 180-200 mm. thick in the middle of the shaft and 120-40 mm. in the belly. The only disadvantage is the increased consumption of cooling water. Construction costs are not higher than are those of the older type, as the decreased amt. of firebrick covers the extra cost of armoring the furnace. B. C. A.

Charging installations of blast furnaces and their effect on the behavior of the run. HUBERT HOFF. *Rev. metal.* 27, 604-14(1930). H. S. VAN KLOOSTER.

Metallography. P. DEVOSSIS. *Camera (Luzern)* 9, 134(1930)—D describes the service of photomicrography in the study of metals and alloys. The prep. of the sample and the arrangement of the microscope are briefly indicated. M. W. S.

Applications of x-rays to the study of metals. JAMES B. FRIAUPE. *Rev. Sci. Instruments* 1, 261-96(1930)—A general popular account is given including the subjects of radiography, metal crystals and alloys, orientation, grain size, internal strain, thermal expansion, fiber structure and the effects of mech. and heat treatment. A bibliography of 44 references is appended. Notice is given of the publication of a complete bibliography on x ray studies of ferrous materials involving more than 300 references. GEORGE L. CLARK.

The recrystallization of metals. P. BECK. *Technika* 11, 151-3(1930)—In case of homogeneous deformation, true grain no./temp. or grain-size/temp. diagrams can be obtained if (1) not only the geometric but also the phys. characteristics of deformation (ϵ , ϵ_c , the original structure and temp.) remain unchanged during the expt., (2) the tested bodies are heated quickly to the resp. temps. and (3) heating is stopped immediately at the end of recrystn. S. S. DE FINALLY.

Microscopic investigation of metals. V. N. KRIVOBOK. *Proc. Eng. Soc. West Penn.* 47, 45-62(1931)—A non-technical review is given emphasizing the purposes of metallurgy and the specific uses of metallography related thereto. With regard to inspection of metals and alloys, microscopic studies reveal the characteristics, the previous history of the sample and its suitability for a given purpose, but only when the samples are properly prep. by an experienced operator. This is illustrated by 7 photomicrographs. The second phase of the metallographist's work is the inspection and investigation of material after a failure has occurred. Under the third category of work, K. discusses the development of new materials, and the illustrations show how to recognize definite changes in properties. The fourth function dealt with is true research or scientific metallography, under which 5 illustrations of martensite decomposition are discussed. M. F. BEHAR.

Plasticity of metals. ZAY JEFFRIES. *Mech. Eng.* 53, 252-6(1931)—Plasticity is defined as "the quality by virtue of which a substance may undergo a permanent change

in shape without rupture." Metals are *cryst.*, and plastic flow is the result of the movement of one position of a crystal with reference to another along crystallographic planes. Industrial metals are useful because of their high plasticity, without which metal-working processes could be not carried on. M F BEHAAS

Gases in metals. C J SMITHFELLS *Metal Ind* (London) 38, 261-4, 268 (1931). —The soly of gases in molten metals usually becomes greater at higher temps., and is generally much less in solid than in liquid metals. With H_2 , N_2 , Pb , Fe , Co , Cu and Pd for alloy systems, alk. earths and alkalis metals form hydrides, and Au , Ag , Be and Mg show practically no absorption. Comparatively few metals absorb N as a gas. The total vol. of gas that can be extd. from brass, Cu or duralumin is of the same order as the volume of metal. Removal of gases from metals involves melting in vacuum, or with Al alloys slow pre-solidification with or without bubbling dry N or Cl gas through the metal, followed by rapid remelting to pouring temps. P R KOSTING

Recent metallurgical research in relation to marine engineering. S L ARCHIBUTT *Trans Inst Marine Eng* 42, 295-346 (1930). —The limiting creep stress should be considered for design purposes at elevated temp. The nature of the atm. seems to influence the rate of creep and the mode of deformation. *Inter-cryst fracture* was found in some creep tests. A graph showing the relation between temp., tensile strength, limiting creep stress and limit of proportionality for 0.17% C steel is given. Higher pressures and degrees of super heat require a material superior to C steel. Creep characteristics are given for various alloys. *Ni-Cr alloys* and their deriv. are most suitable for temp. ranges of 600-1000°. Limiting creep stress at the highest temp. increases with increasing Cr content. ANN NICHOLSON HIRD

Use of non-ferrous metals in the aeronautical industry. D HANSON *J Inst Metals*, Adv. copy, No 545, 24 pp (1930). —The alloys of Al and of Mg now used in the aeronautical industry are considered. Mg is inferior in comparison with Al in that it has a lower modulus of elasticity and is not as resistant to corrosion. J L GREGG

Magnetic method for testing of wire ropes, especially wire cable. FRANZ WEYER AND ARNOLD OTTO *Mitt Kaiser-Wilhelm Inst Essenforsch.* Dusseldorf 12, 389-90 (1930). —The device developed for boiler tubes was investigated as to its suitability for cables. It proved to be sufficiently sensitive to permit the certain detns. of single wire fractures even in the interior of the cable. R. RIMBACH

Pickling and burning. L. HERBERT KURRIEN *Chem-Ztg* 55, 133-4 (1931). —Pickling serves to remove foreign material from metal surfaces. The mechanism of the process consists of 3 stages: loosening of the scale, soln. of the scale and, to some extent, soln. of the metal. Scale is loosened by H evolved from the metal. H_2SO_4 is preferable to HCl in pickling, because it may be heated without increasing the vol. of objectionable fumes. Some portions of the scale are more tenacious than others and require longer pickling. Unless inhibitors are added, the clean metal will be dissolved, and diffusion of H from the reaction will embrittle the metal. Inhibitors should not materially increase pickling time or remain on the metal surface. The metal concn. in the pickle should not be allowed to become higher than 50-60 g. per l. After pickling, the metal should be rinsed to remove reaction products and the excess acid neutralized with dil. Na_2CO_3 to prevent fresh attack on the metal. *IL Ibid* 150-1. —"Burning" is the process of treating metals with strong mixed acids for the purpose of obtaining dull, colored or shiny finishes. Various formulas are given for "burning" solns. The acids used are H_2SO_4 and HNO_3 in varying proportions. $NaCl$, $K_2Cr_2O_7$, and H_2F_2 and various forms of lampblack are sometimes added to obtain different finishes. Precautions for operating are discussed. B E ROPTHLETT

Metal-ware pickling data. WALLACE G. IMHOFF *Am Metal Market* 38, No 54, 5, 10 (Mar 20, 1931), No 55, 5 (Mar 21, 1931), No 56, 5, 10 (Mar 24, 1931). —Data included in 28 tables showing the results in a no. of plants of the pickling of hot galvanizing of 12-qt. light water pails are discussed. Details are given for wts. of pails, pickled wts., the strength of pickles used, etc. The wts. of Zn deposited differed by 19½ lbs. per gross deposited as coating. W H BOYNTON

Modern case-hardening practice. FRANCIS W. ROWE *J West Scot Iron Steel Inst* 38, 27-42 (1930). —The correct choice of steel for case hardening depends on the character of the stresses to which the part will subsequently be subjected in service. These steels are C steel, 3% Ni , 3.5% Ni , 3.5% $Ni-Cr$, 5% Ni , 4% $Ni-1\%$ Cr , $Ni-Mo$ and $Cr-Mo$ steels. Plain carbon steels are subjected to abnormality. Ni case hardening steels give higher yield point and tensile strength than plain C steels. The presence of 3% Ni enables a satisfactory hardness to be developed by a final quench in oil, and the case is rendered tougher. In order to insure proper hardening of the core the steel

must contain 0.13% C and 3-4% Ni, and be quenched from 780°. A class of case hardening steel which is becoming increasingly popular is that with 4% Ni and 1% Cr, with 0.16% C and with 0.20-0.30% Mo. This material in the water-quenched condition has a strength 90-95 tons/sq in and 80-85 tons/sq in in the oil-quenched condition. This is associated with an impact value of about 30 ft. lbs. Such steel is capable of withstanding the highest surface stresses, even with a comparatively shallow case, and is used for gear-box gears. The disadvantage of this steel is that, even in the fully annealed condition it is relatively hard to machine. It is also much more prone to over-carburizing than the lower alloy-content steels, and for this reason the carburizing temp. should be low. G. T. MORON.

The nitriding process of nitrogen case hardening. J. I. HURAR. *Iron & Steel Ind. Brit. Foundryman* 3, 351-4, 370-80 (1930).—The steels used in nitriding and the details of this process are discussed. LEOPOLD PESSER.

Properties of meehanite metal—a pearlitic iron. A. G. LAMBERT AND F. M. ROBBINS. *Can. Foundryman* 21, No. 9, 14-5 (1930).—Data are given on comp. and phys. properties. E. I. S.

Influence of the emissive power on the temperature measurement of liquid iron. RUDOLF HARR. *Arch. Eisenhüttenw.* 4, 261-4 (1930).—The influence of oxidation on the emissive power of liquid Fe, and therefore on optical temp. measurements, was studied by optical, thermoelec. and photographic methods. Optical measurements made on oxidized surfaces agree more nearly with thermoelec. measurements than those made on unoxidized surfaces. The latter may vary as much as 80-140°. Between 1250° and 1600° the radiating power of bare iron is 0.44 ± 0.03 , while that of oxidized iron is 0.65 ± 0.05 . These values are practically independent of temp. Deoxidation due to the dissolving of the oxide begins at 1400° and increases with temp. until at 1580° an oxide coating is no longer produced by air. H. F. JOHNSTON.

Results of operation and fields of application of various types of malleable iron recently produced. RUD. STORZ. *Die Giesserei* 18, 1-8 (1931).—A lecture. J. H.

High-quality malleable iron. E. PIKOWARSKY. *Die Giesserei* 18, 19-24 (1931), cf. *C. A.* 24, 5204.—A lecture. J. BALOGIAN.

The influence of nickel and nickel chromium on the properties of malleable iron. LEON TRIEKY. *Rev. metal.* 28, 1-18, 61-78 (1931), cf. *C. A.* 24, 4744.—White irons prep'd to contain C 3, Si 0.4, Mn 0.10-0.20 and S 0.018-0.009% were alloyed with 0.5, 1.4 and 2.5% Ni. They were tested together with the original base and an alloy produced by alloying it with 1.37% Ni and 0.34% Cr. Specimens were cast in sand but removed from it rapidly enough to produce chill. On plain C stock, the dilation curves show the effect of quenching. The decompn. of cementite in them begins at 1100°. On repeated heating, the cementite begins to decompose at Ac. The prtn. of graphite on the first heating can be seen under the microscope. With 0.5% Ni, the dilation curves have the same appearance, at least on the first heating, but the cementite begins to decompose at Ac. The presence of 1.4% Ni lowers, by 100°, the temp. of graphitization, reduces the temp. of γ to α transformation and lowers, by 50°, the Ar point. Addn. of 2.5% Ni considerably increases the quenching action of the Fe, reduces Ac below 700° and causes graphitization at 870°. The addn. of 1.37% Ni and 0.34% Cr increases the cementite anomaly and raises the temp. of graphitization, reducing the tendency toward quenching by raising the Ar point. Malleabilizing treatments showed that Ni helps to stabilize the system by decomp. the cementite. It prevents changes on cooling and decreases the size of the constituents proportionally to its content. Cr exerts an opposite action. From 1 to 2% Ni increases the tensile strength by about 30% and almost doubles the elongation. Toughness also increases, reaching 0.430 kg./m. for 1.5% Ni. J. D. GAT.

Accelerated annealing of malleable cast iron. V. V. USOV AND N. N. PUSTUMIKOV. *Vestnik Metalloprod.* 10, 97-103 (1930), *Khimie & Industrie* 25, 352 (1931).—The American process is too long, requiring 100-120 hrs. The General Electric Co. has succeeded in reducing the time required to 28-30 hrs., but this requires the use of an elec. furnace. U and P. have endeavored to render the process suitable for application in ordinary case-hardening furnaces. Their work has shown that (1) In order to accelerate graphitization of the free cementite, it is merely necessary to keep the iron above 1000° for 2 hrs., at the end of which time the metal consists exclusively of the solid soln. and of graphite, (2) To reduce the time required for graphitization of the cementite which seps. on cooling, the metal must be cooled rapidly to 710° and maintained 4 hrs. at this temp. to decompose the pearlite formed by cooling into ferrite and graphite. This process required only 8-10 hrs. in the lab. and 16 hrs. in a petroleum-heated muffle furnace. A. PAPINEAU-COUTURE.

Graphitization in cast iron. OTTHAR V KFIL *Arch Eisenhüttenw* 4, 245-50 (1930) —The development of graphite in cast Fe depends on the position of the eutectic. Carbide congelation at the min of the arrest point shows very fine granular graphite sepn, while at the max needle-like formation occurs. The cooling velocity before the eutectic is retarded by increasing C content because of the decreasing work of diffusion. The transition of the carbide forms after the stable congelation has a crit velocity dependent upon the C content, a fact which indicates a mol transformation to a stable phase. Overheating reduces the crit cooling velocity. Overheated cast Fe congeals throughout in the metastable form and remains in this condition after remelting.

H F JOHNSTONE

Nickel cast iron and its uses in the electrical industry. W C HIRSCH *Elec Mfg* 5, No 2, 55-7 (1930) —Marked improvement obtained in the quality of steels by addn of Ni led to an intensive investigation of benefits obtainable from its use in gray iron. Ni cast iron offers a solution to many problems in the manuf of elec. machinery and domestic appliances.

E I S

The theory of the tempering of cast iron according to American methods. N HEKKER. *Die Giesserei* 18, 14-9, 39-45 (1931) —A theoretical paper. J B.

Fatigue tests on iron and steel. AXEL LUNDGREN *Jernkontorets Ann* 115, 1-70 (1931) —Besides a detailed review of the most important results obtained in different countries by studies on the fatigue of metals, a report is given on investigations carried out at the Government Testing Inst., Stockholm. The results of these investigations, which were made with the 'Alpha' testing machine, may be summarized as follows: Test bars for one-sided load were made with radius of fillet 0, 1, 2, 5 and 10 mm. A radius of 10 mm gave 90% higher fatigue limit than a radius of 0 mm and 2% higher than a radius of 5 mm. Tests with const moment gave a 2-3% higher fatigue limit than tests with one-sided load. A number of steels were investigated with the object of studying the relation between endurance and heat treatment, as well as between endurance and other mech properties. Widely differing annealing temps were found to have a very slight influence on the fatigue limit. On the other hand the endurance is considerably increased by quenching and varies with the degree of tempering. No marked relation was found between the fatigue limit and the limit of proportionality or the yield point. The relation between endurance and ultimate stress is more pronounced. A scrutiny of the results of this and previous investigations indicated that the ratio K of fatigue limit to ultimate stress varies within certain limits and increases with reduction of area. By plotting K against reduction, diagrams were obtained for steels with different microstructure, from which the following conclusions may be drawn: *Pearlitic C steels* — $K = 0.37$ for a reduction of area less than 36%. Above this value K increases with reduction. *Pearlitic Ni and Cr steels* —Values of K than in pearlitic C steels. *Sorbic steels* —Values of K are considerably higher than in pearlitic steels. Increase of K with reduction of area is not so pronounced as in pearlitic steels. *Steels conig a mixt of sorbite, pearlite and ferrite*, with the ferrite in the form of network. K is considerably lower than in sorbitic steels, sometimes even lower than in pearlitic. *Martensitic-baustitic steels* with high ultimate stress. Values of K are lower than in sorbitic steels. A one-page English summary is given. (43 references)

H. C DURS

The influence of various elements on the carburization of iron and steel. GENSKÉ TAEHASHI *Kinzoku no Kenkyu (J for Study of Metals)* 8, 102-26 (1931) —The influence of various elements on the degree of carburization of special iron and steel and also the same influence on the hardness of carburized products have been systematically investigated.

M KURODA

Optical investigations on the passivity of iron and steel. L TROSTAD *Nature* 127, 127-8 (1931), cf *C A* 24, 1566 —Using mirrors of Fe and steel, T treated them electrolytically in baths of alk., neutral and acid Na_2SO_4 solns., different c. ds were used. During the passive state all mirrors showed a change in the reflected light, this corresponded to the formation of a surface film with a mean refractive index of 3.0 and a mean thickness of 30 Å. This agrees with a film of Fe_2O_3 . Reactivation did not completely remove the oxide skin. More of the film was destroyed on ferrite than on cementite grains. Alternating cathodic and anodic treatments 3 times in 3 hrs left microscopically visible dark brown spots on the cementite, and faint yellow interference colors could be detected only by employing slanting incidence over the ferrite particles.

FRANK MARESH

Intercrystalline fractures in soft steel. ENRICO CREPAZ *Atti III congresso naz chim. pura applicata* 1930, 380-7 —An attempt was made to explain the formation of intercryst cracks in soft steel (0.14% C) after immersion of the steel in liquid metals

or alloys, such as Cu, Zn, brass (30% Zn), bronze (15% Sn) and Al bronze (7% Al). Steel rods of 1 cm diam were immersed for 2-3 min. in the molten bath and were bent after removal from the melt (1) when the adhering melt had not yet solidified and (2) after complete solidification of the molten alloy. In both cases transverse cracks appear which are deepest at the places where the curvature of the rod is greatest. Microscopic observation of transverse sections showed that the cracks were intercryst. and filled with the solidified metal or alloy. The cracks are due to the internal stresses set up as a result of localized heating of the steel in contact with the liquid alloy. This is also brought out by the fact that bending of the rods inside the bath after immersion for several min. does not, as a rule, produce cracking, whereas bending immediately after immersion causes, in most cases, pronounced intercryst. cracks. The dissolving action of the melt is not a contributing factor, since the fractures appear more readily in brass and bronze than in Zn, although the latter has greater dissolving power.

H. S. VAN KLOOSTRA

Mechanism of coalescence in steels. R. K. THAKURE *Proc. 15th Indian Sci. Cong.* 1928, 92.—In considering the phenomenon of coalescence in alloys with special reference to coalescence of carbide in steel from the viewpoint of theoretical physics, K. makes 2 assumptions: (1) slight sol. of carbide in ferrite below the A_1 point, and (2) diffusion in solid state subsequent to state of inequilibrium. Both are well supported by the well known researches of Benedicks, Roberts-Austen and others. K.'s observations (not given) show that the acceleration in coalescence in steels due to cold working is not primarily due to breaking up of cementite lamellae. K. made, but does not include here, a quant. record of progress of coalescence of carbide in sorbitic steels, as measured by degradation of mech. properties of steel.

E. J. C.

New process of quenching steel. N. F. BOLKHOVITSOV *Vestnik Metalloprod.* 10, 88-90 (1930), *Chimie & Industrie* 25, 352-3 (1931).—Is exptd. with Lewis' process, which consists in heating the steel above the crit. temp. and then quenching in a bath of molten salt or of oil heated to 230°. When treated in this way, the steel is at first soft, it consists of austenite, and becomes hard only about 3 min. after having reached ordinary temp. It then expands, becomes magnetic and is transformed into martensite. Owing to this slow change in structure, all the transformations produced by the variations in vol. have time to take place, which completely eliminates the usual defects of the ordinary quenching process, giving great hardness and homogeneity. But from a practical standpoint, the greatest advantage of the method lies in the fact that pieces quenched in this way can undergo all the requisite deformations and flexions in 1-2 min. after quenching.

A. PAPINEAU COUTURA

Interpretation of the complex phenomena observed during the reversion of highly tempered steel. ALBERT PORTAVIN AND PIERRE CHEVENARD *Compt. rend.* 191, 108-10 (1930).—Two reactions take place simultaneously: (1) Supersatd γ -iron decomps into cementite and α -iron, with less C causing contraction and (2) γ soln decomps into cementite and α -soln with little C causing expansion. Any change in dilatometric manifestation is due to a preponderance of one reaction or the other, according to the temp.

O. M. SMITH

The resistance of steel to cyclic stresses. W. SCHNEIDER *Stahl u. Eisen* 51, 285-92 (1931).—The importance of the resistance of steel to cyclic stresses in construction work is pointed out. The difference between the long- and short-time testing methods and the effect of surface irregularities and of combined stresses on the results of the tests are discussed.

J. A. SZILARD

Annealing of steel in a protective atmosphere of producer gas. A. JOHANSSON AND E. VON WACHENFELDT *Jernkontorets Ann.* June 1929, 141-79, 180-9.—Results of an investigation of the possibility of bright-annealing steel with charcoal producer gas show that this can be done. The app. used is described.

E. I. S.

Coarse-grained recrystallization of cold-drawn low-carbon seamless tubes. Annealing tests with critically cold-rolled rods of low-carbon steel. ANTON POMER AND ERWIN HOLWEG *Miss Kaiser-Wilhelm Inst. Eisenforsch. Dusseldorf* 13, 1-23 (1930).—It was detd. that the grain size of critically cold-rolled rods was considerably smaller after annealing at 750° than after annealing at about 700° or about 800°.

R. R.

Effect of various annealing temperatures on cold-worked low-carbon steel. H. E. PUBLOW, C. M. HEATH AND R. A. GAZELUS. *Mech. Eng. Expt. Sta., Bull. No. 25*, 17 pp. (1931), cf. *C. A.* 24, 3975.—Without appreciable change in microstructure, low C steel (analysis not given) changes in hardness from 58 to 85 Rockwell B on drawing into brake drums. Olsen ductility tests for such a steel give 0.17 in. deflection at 9000 lb./sq. in. max. load. Annealing at 1150-1250°F for 1/2 hr. and cooling at the rate of 70°F per min. dropped the hardness to 60 and raised the deflection to 0.36 in., max.

load to over 11,000 lb./sq. in. This gave better ductility than raising above the crit. point and cooling at that rate. Various other times and temps were tried. Much of the hardness produced by drawing into brake drums is acquired before the microstructure shows apparent distortion of the grains. Annealing at 800°F increases the hardness instead of decreasing it, prolonged heating below 1000°F does not relieve internal strains, the best annealing temp is 1250–1300°F, *s e.*, just below the lower crit. point.

H. W. GILLET

The sensitiveness to overheating of low-carbon ingot steel. E. POHL, E. KRIEGER AND F. SAUERWALD. *Stahl u. Eisen* 51, 321–6 (1931).—Annealing tests between 950° and 1400°, in 50° steps, were carried out on 4 low C ingot steel samples and one Mo steel. The temp at which oxidation occurs after 1 and 5 hrs. of annealing, the depth of the oxidized layer and the change in grain size as a function of the annealing temp were detd. After 1 hr. annealing the temp. of overheating was found to be between 1200° and 1400° for the various samples and the thickness of the oxidized layer 0.1–0.2 mm. After 5 hrs. annealing the temp. of overheating was found to be between 1150° and 1200° and the depth of the oxidized layer 0.4–1.4 mm. Ingot steel with a coarse grain structure is less resistant to overheating than steel with a fine grain structure.

J. A. SZILARD

Copper steel for dynamo and transformer sheets. A. KESSELMANN, B. SCHAANOW AND V. S. MESSIN. *Stahl u. Eisen* 50, 1154–7 (1930).—Addn. of up to 0.7% Cu to steel with 1.5 and 4% Si has practically no effect on the magnetic properties, but a larger proportion of Cu has a deleterious effect. Less than 0.7% Cu improves the hardness and tensile strength and the resistance to corrosion of these steels and is therefore a desirable addn.

B. C. A.

Fatigue strength of carbon- and alloy-steel plates as used for laminated springs. R. G. C. BATSON AND J. BRADLEY. *Engineering* 131, 405–6 (1931). Conditions of the surface layer of spring-steel plates cause fatigue resistance of complete laminated springs to be lower than the mech. properties of material would indicate. Fractures of laminated springs may be due to very low limiting ranges of stress obtained from unmachined plates compared with those from machined specimens. Deleterious layers could be removed by machining $\frac{1}{4}$ in. from the surface. Surface effect was found to be produced by hardening and tempering. Only very slight improvement was obtained by heat treating spring plates after a thin layer had been machined from the surface of the rolled material. Tables are given summarizing the results of tests made.

Lois F. McCORMACK

Cobalt magnet steels. H. E. KEASHAW. *Edgar Allen News* 9, 701–5 (1930), *Metals & Alloys* 2, Abstracts 63.—There are 2 classes of Co magnet steels, the low- and medium-Co chrome steels and the high-Co steels. The Co chrome steels are subdivided, one contg. 9% and the other 16% Co, but both contg. Cr 9–10, Mo 1.5 and C 1.0–1.15%, these are air hardening, and the best results follow a triple heat treatment. The high-Co steel usually has Co 36, Cr about 6, W 4–6 and C 0.9%—these are oil hardening, only a single treatment is necessary, but accurate temp. control is essential. This alloy has better magnetic properties than any other known at present, *B H* max. values of 900,000–1,600,000 are regularly obtained. The high frequency elec. crucible furnace is very efficient for the manuf. of Co magnet steels and can also be used for cast magnets. With Co-Cr steels better magnetic results are obtained from forged than from cast magnets—the higher the Co content, the greater the discrepancy between the magnetic results. With the 36% Co alloy, cast magnets can be produced having as good, sometimes better, magnetic properties than forged magnets. Some special magnets are described.

G. G.

Properties and application of nickel and nickel-chromium steels. J. GALIBOURG. *Pea metal* 28, 30–9, 85–100 (1931).—A review of the properties of plain C, Ni, Ni-Cr and stainless Ni-Cr alloys.

J. D. GAT

Suppressed constitutional changes in alloys. G. SACHS. *Am. Inst. Mining Met. Eng., Preprint Inst. Metals Division* 1931, N. Y. meeting, 10 pp.—X ray analysis and single-crystal study have been utilized in recent years as a new means of following constitutional changes in alloys. If such transformations can be suppressed by rapid cooling, they can be followed at room temp. in a particularly convenient manner. This paper considers changes in soly and polymorphic transformations with particular reference to phenomena of the age hardening type. A thermodynamic viewpoint is developed for examg. such transformations.

C. L. MANTELL

The structure of cast alloys. MARIE L. V. GAYLER. *Metallwirtschaft* 10, 141–4 (1931).—An Al alloy contg. 7% Cu was melted in a high frequency induction furnace and held in a vacuum, in the presence of air, N, H and furnace gases for 25 min. One

set was cast at 740° and another at 840°. Each heat was cast into 4 different molds: steel, graphite, water cooled Cu and sand. Photomicrographs of the structures are shown. The furnace atm. has little effect on the specimens cast at 740°. The macrostructure in all is fine grained, in those melted in vacuum and under H it is the finest. Those cast in sand are somewhat coarser than those cast in steel. The microstructure is the same in all. The macrostructure of the set cast at 840° into steel molds is coarser than that cast at 740°, and there is much more difference between different furnace atms. Furnace gases produce the coarsest grain, and the difference is still greater when cast in sand. The microstructure is finer grained, at the higher casting temp. Both series cast in water cooled molds have a similar fine grained macrostructure, with little difference between the various furnace atms. C. E. MCPARLANA

An investigation of wrought light alloys. HUNTARÖ ÖTAVI. *Kansoku no Kenkyu (J for Study of Metals)* 8, 89-101 (1931).—A relative test of forgeability for 40 kinds of special light alloys was carried out, and the effect of heat treatment on the hardness of these alloys was also investigated. For the alloys which showed good mech. and working properties, the effect of natural and artificial aging on the hardness of quenched alloys was investigated, several mech. tests were also carried out on the bars and plates of these alloys. M. KUROO

Production and application of the light metals and of their alloys in Italy. A. W. BONARETTI. *Metalli leggeri* 1, 7-17 (1931).—A monograph. ROBERT S. POSMONTIER

Aluminum and its alloys. P. LIBERA. *Metalli leggeri* 1, 19-28 (1931).—A description with photographs. ROBERT S. POSMONTIER

Influence of antimony in aluminum alloys. A. W. BONARETTI. *Metalli leggeri* 29-52 (1931).—A description is given of the phys. and chem. properties and compns. of various Al alloys contg. Sb, especially treating of the resistance of such alloys to action by sea water. Many photographs are given. ROBERT S. POSMONTIER

Changes in properties during the cooling of supersaturated silver-copper alloys. N. V. AGEV, M. HANSEN AND G. SACHS. *Z. Physik* 66, 350-76 (1930).—In general the changes proceed in the same manner whether the alloy is rich in Ag or in Cu. In addn. to the normal changes in resistance due to the altered compn. there is an anomalous effect producing an increased resistance. At low concn. and at high temp., the change in hardness and the anomalous resistance are related to the amt. of sepd. crystals. At high concn. and low temp. there is a marked increase in hardness in the initial stages of sepn. without any corresponding change in resistance. S. BRADFORD STONE

The structure of copper-zinc alloys. W. BRONIEWSKI AND J. STRASBURGER. *Rev. metal* 28, 10-29, 79-84 (1931), cf. *C. A.* 24, 4747.—Literary data do not definitely establish the compn. of the definite compds. of the Cu-Zn system. To obtain the necessary equl. specimens cast in chills and contg. less than 49% Zn were annealed at 400° for 1000 hrs., and those with higher content of Zn for 3000 hrs. The treatment was conducted in an elec. furnace well luted with fireclay, so that the changes in the compn. were found only in a very thin outside layer of specimens. Parallel series of tests were made on cast and so-annealed specimens. Elec. cond., thermoelectric power, ϵ in f of the soln., dilation and Brinell hardness were detd. These tests and microscopic evidence placed the limits of solid soly. of the α phase between 0 and 35 at % of β between 46 and 49%, of γ between 59 and 67%. After 3000 hrs. annealing the β phase was never perfectly homogeneous, this interfered with the detn. of its limits of solid soly. which tentatively can be given as 79 and 87%. The range for α phase is within 94 and 100% Zn. The existence of CuZn , CuZn_2 and CuZn_3 at room temps. was reasonably well established. The shifting of the maxima of corresponding curves from theoretical 50% to about 49.6% can be explained on the assumption of the partial decompn. of CuZn into γ - and α phase. J. D. GAT

Platinum-iridium alloys. FRANZ KORN. *Metal Ind.* (London) 38, 309-10 (1931).—Pt-Ir alloys have high strength and a faint blue color and are resistant to tarnishing. The tensile strength varies from 30 tons/sq. in. with 5% Ir to 97 tons/sq. in. with 30% Ir. Alloys contg. more than 12% Ir increase in tensile strength when annealed between 600° and 800°. Annealing temps. should never be above 950° especially for alloys with more than 15% Ir, some of which is lost by volatilization. Corrosion resistance to hot aqua regia and H_2SO_4 increases with Ir addns. Pd does not harden Pt alloys as effectively as Ir. Au increases both hardness and brittleness of Pt, but the brittleness is counteracted by Pd. The Pd content should be 2 or 3 times the Au content to obtain the best results with Pt-Pd-Au alloys. P. R. KOSTING

Development of zinc-base die-casting alloys. D. L. COLWELL. *Proc. Am. Soc. Testing Materials* 30, Pt. 11, 473-92 (1930).—The history of the improvement of Zn base die casting alloys is illustrated in series of charts, which show the effects of nor-

mal and accelerated aging on size and phys properties of Zn base die-castings. The data from the tests and the conclusions are given. H. W. GILLET

Equilibrium diagram of the system: lead-arsenic. MOHAMMAD OMAR FARUQ *Proc 15th Indian Sci Cong* 1928, 176—Arsenic alloys with Pb in all proportions. The diagram was studied up to 60°C. As—Pb does not dissolve any As at its m. p., but at higher temps 2 layers are formed. The lower layer is Pb contg a small quantity of As in it, and the upper layer is As with a small quantity of Pb dissolved in it. W. J. C.

Thermal analysis of the system: lithium-copper. S. PASTORELLO *Gazz chim ital* 60, 988-92 (1930). Since the publication of a paper on Li-Ag alloys (cf. C. A. 25, 270) a method for the thermal analysis of such systems has been developed, which requires only 0.3-2.0 g. of Li and permits verifying a eutectic almost to the limit of its formation. A systematic series of thermal analyses of binary alloys of Li and heavy metals of the same group was therefore begun. There is no earlier literature on Li-Cu alloys and the constitutional diagram of such a binary system is unknown. Preliminary to the thermal analysis of Li-Cu an x-ray examn was made by photographing an alloy contg approx 50% of Li. This gave a roentgenograph with only diffraction lines which correspond to the normal lattices of Cu and of Li. Consequently no intramol compd and no solid soln of Li with Cu were formed. This behavior, which was then confirmed by thermal analysis, is quite different from the behavior of Li with Ag as already shown (cf. C. A. 25, 270) and of Li with Au as x-ray examn, to be described later, shows. The Li-Cu alloys (1.7 to 94.01 at % of Li) were prepd by fusion in a Ni-Cr steel crucible in an A atm, and the cooling curve was obtained by temp measurements with a Pt-Pt-Rh thermocouple and a Hartmann and Braun galvanometer. If special Ni-Cr steel is not used, the mixt attacks the crucible badly, particularly with alloys rich in Cu. Even with Ni-Cr steel, several preliminary fusions must be carried out until the % Fe in the alloys has become low enough (Cu contg 3% Fe is only 10% different from pure Cu—cf. Ruer and Goerens, C. A. 12, 577; Sahmen, C. A. 2, 154)—so 0.1% Fe in the Li-Cu alloy is low enough. Alloys contg 1.71, 9.90 and 14.38 at % Li had a color resembling pure Cu, whereas with higher and increasing % Li the color approached the metal gray of Li. All alloys were ductile, and those with high % Li were almost as soft as pure Li. All were readily attacked by water, leaving cryst. Cu of unusual appearance with alloys high in Li. A eutectic was formed at the m. p. of pure Li, even for alloys contg less than 2 at % Li. In this latter case the m. p. was approx 50° below that of Cu. From 10 to 90 at % Li, the m. p. depression was almost linear. The 2nd arrest in the m. p. of pure Li was found with all the alloys. Pure Li melts at 180°, which agrees with the data of the majority of investigators. Alloys rich in Li did not show any marked liquidation, as would be expected from the great difference in d. C. C. DAVIS

The solid solutions of the copper-silver system. D. STOCKDALE *J Inst Metals*, Advance copy No. 359, 14 pp (1931).—The solid solubilities of Ag in Cu and Cu in Ag were detd by microscopic and elec methods. The solv of Ag in Cu is 8.2% at the eutectic temp., 4.7% at 700°, 0.7% at 400° and probably almost zero at room temp. The solid solv of Cu in Ag is 8.8% at the eutectic temp., 5.8% at 700°, 1.1% at 400° and about 1% at room temp. The Cu-rich solidus curves downward sharply near the limit of solid solv. The Ag-rich solidus is nearly a straight line. J. L. GREGG

The effect of rate of bending in notched-bar bending tests. JAMES G. DOCHERTY. *Engineering* 131, 347-50 (1931).—The expts. were undertaken to establish the relation between the slow bend and the Izod notched bar test, and to discover whether the results obtained in the Izod test measure the resistance of the material to impactive loading or to the stress concn produced by the notch. A special testing machine, devised by D., was used. The materials tested were C steels, 3% Ni steel, Monel metal, naval brass, phosphor bronze. In the more brittle metals, i. e., those with Izod impact strength less than 30 ft. lb., it has been possible to test at 1/2 the Izod speed in the standard Izod machine by reducing the available energy to 50 ft. lb. These tests help to bridge the unexplored gap between the fastest "slow bend" test (150 in./min.) and the standard Izod test (7000/min.). For all the ductile metals tested the energy absorbed in bending or in fracture increases with the speed of test. Impact brittleness is considered as a velocity effect, and there is no discontinuity between the slow bend and Izod test. *Ibid* 414-5.—The relation between the notched bar test and the ordinary tensile test is energy absorbed in ft. lb. divided by ultimate tensile strength in tons per sq. in. G. T. MOROX

Measurements on the degree of orientation in hard-drawn copper wires. W. A. WOOD *Phil Mag* 11, 610-7 (1931).—The results confirm Schmid's (cf. C. A. 23,

1029) The degree of orientation was estd by the measurement of photographic film intensity at convenient points on the diffraction rings where max or min occur. The drawn wire may be described as consisting of a randomly oriented cylinder surrounding a highly oriented core. This is characteristic of all drawn wire. The effect of drawing is to slowly increase the degree of orientation in the interior of the wire. Orientation is highest along the axis of the wire. The change from oriented to non-oriented is quite sharp.

ARTHUR FLEISCHER

Cold-working and annealing of α brass. J. NAVARRO ALCÁZAR. *Anales soc espn fis quim* 28, 1429-31 (1930).—The process of recrystn. of cold-worked brass was studied, and the optimum annealing temp. detd. There were detd. the x ray diagrams, which fixed the positions of the elementary lattices, the temp. of beginning of recrystn., Lane grams and the progressive increase of grain at different temps. Mech. tests gave data on improvement of coeffs. of cold working by repeated annealing and recrystn. processes which agreed well with the above. In stamping and the like the best results are obtained between 800° and 900°.

E. M. SYRAGES

The recrystallization of technical lead. F. LOOPS-BASSOW. *Metallwirtschaft* 10, 161-5 (1931).—The grain sizes of 17 grades of Pb were detd. after casting them into strips, reducing them 50% by rolling and annealing for 15 hrs. at 200°. Electrolytic Pb has a grain size considerably above 1 mm., com. Parkes Pb about 1 mm., and Pattison Pb and Pb alloys less than 1 mm. Elements like Cu and Ni which are not sol. in Pb reduce its grain size, while Bi, which is sol., does not. Samples of one of each of the 3 grades of Pb were reduced 6-8% and annealed at 18-310°. Grain size was detd. and recrystn. diagrams constructed. The increase in grain size with rising annealing temp. is gradual for the first 2 grades, but Pattison Pb hardly changes up to 280°, then suddenly becomes larger than the other 2. The lower recrystn. limit is above room temp. for very low deformations, below for higher deformations. There was no difference in the recrystn. grain size of samples rolled at room temp. and at -70°. Long aging of rolled samples at room temp. produces normal recrystn. with higher deformations and grain recovery with low deformation.

C. E. McFARLANE

The influence of impurities in soft lead on its behavior toward acids. ARTHUR BURKHARDT. *Metallwirtschaft* 10, 181-7 (1931).—Forty-one com. grades and alloys of Pb contg. small percentages of Cu, Ni, Bi, Sn, Sb, Te, Zn, Cd, Ag, Ti, Fe, Hg or Li, or combinations of these were tested for corrosion resistance against various strengths of H_2SO_4 and HCl at 20°, 90°, 130°, 170° and 220°. Tensile tests are not a suitable measure, as the strength drops in proportion to the reduction in cross sections of the test bar. Loss of weight was detd. instead. Seventy % H_2SO_4 at 170° attacks com. Pb contg. Bi, Zn or Ag most, then those contg. little or no Cu, and those contg. Ni or Cu the least. Cu can compensate for the bad influence of Bi. At 90° Sn, Sb and Ag are no worse than Cu alloys, and at 20° there is very little difference between any of the grades. With 98% H_2SO_4 similar results were obtained. Fe, Sb, Sn, Te, Cu and Ni retard corrosion of Pb in H_2SO_4 . Ti and Hg have no effect, Li, Ag, Zn, Cd and Bi increase corrosion. Tests were also made in 10 and 20% HCl and in moist Cl_2 at 20° and 90°. In HCl brands which contain Cu, Ni and Te are attacked most and the purest grades least. In 20% HCl the attack is greater than in H_2SO_4 , but in 10% HCl it is slight, in moist Cl_2 it is a little greater. Attempts to obtain greater resistance to HCl by alloying failed. No theory for the effect of impurities on the corrosion of Pb has been worked out, it must be detd. empirically.

C. E. McF.

The oxidation and intercrystalline brittleness of nickel. N. V. AGEEV. *Trans Inst Metals* (Moscow) No. 7, 61-79 (1930).—Two grades of com. Ni showed that Ni is brittle after having been annealed under atm. conditions at 800° and 1000°. Intercryst. oxidation is the cause of this brittleness.

E. I. S.

Sprayed aluminum coatings. LEOPOLD PESSER. *Metal Ind* (London) 38, 288 (1931).—A blister type of corrosion, occurring on rolled Al sheet, is described and theoretically explained. These corrosion exfoliations can be prevented by the application of an Al coating by the metal spraying process.

LEOPOLD PESSER

Standards and exposure tests for plated metals. WILLIAM BLUM. *Brass World* 27, 57-60 (1931).—A series of exposure tests of electroplated surfaces has been started. Points considered include the location of the tests, order of tests, solns. and conditions to be used in plating, size, shape and arrangement of specimens, prepn. for plating, conditions of exposures, methods of inspection, etc. Any general specifications must be prepd. with respect to certain types of use or service and adopted or modified by mutual agreement for any specified article.

W. H. BOYNTON

Estimating galvanizing pot life. WALLACE G. IMHOFF. *Iron Age* 125, 1154-5 (1930), cf. C. A. 24, 2973.—The temp. of the bath is the most important single factor

influencing the destruction of the galvanizing pots. They should be designed in such a manner that no part can possibly be heated above 900°F. The regular operating temp of the pots should be as low as possible and should never go above 875-85°F. The use of an elec. pyrometer for indicating and recording the pot temp is recommended. A table is given showing the relationship between the bath temp. and the loss of thickness of the steel walls, from which deductions as to the estd life of galvanizing pots with a wall thickness of 1 in and 1.25 in are drawn. Practical experience was found to check up with these theoretical figures. LEOPOLD PESSEL.

Galvanizing-pot destruction. WALLACE G IMHOFF *Iron Age* 125, 633-6(1930); cf. C. A. 24, 2973 —The attack of molten Zn upon steel sheets was studied at different temps. The attack becomes very strong at 900°F. LEOPOLD PESSEL.

Comparison of rust protection of iron by zinc, by cadmium and by zinc-copper alloys and the electrodeposition of such alloys. FRANK C MATHERS AND RUSSEL L HARDY. *Proc. Indiana Acad. Sci* 38, 183-5(1929) —Measurements were made of the relative degree of protection against rusting of an Fe surface given by Zn, by Cd and by Zn-Cu alloys. The metals were sprayed with 5% salt soln to induce rapid rusting. The presence of Cu in Zn deposit on Fe lessens the time of protection in the salt spray. Cd gives almost twice as long a protection as does an equal wt. of Zn. Cost of rust protection by Cd is 7 times that by Zn. A method is given for controlling the compn. of electrodeposited alloys by using anodes of 2 different compns and switching the current alternately from one to the other. H M STARK.

Apparatus suitable for control determinations of the total zinc coating on steel, GÖRKE, *Chem. Fabrik* 1931, 147-8 —The vol of H evolved by acid is measured. Cf. Cushman, C. A. 15, 1123, 1882. J. H. MOORE.

Corrosion and its prevention. J VESZELKA *Bányász Kohász Lapok* 63, 9-14, 35-41(1930) —A brief account of corrosion and non-corrosive metals and alloys which may be used in boilers, etc. S. S. DE FINALLY.

The corrosion resistance of copper-bearing steels. W MARZAITN AND A PUSCH *Korrosion Metallschutz* 7, No 2, 34-9(1931) —Lab tests of Cu-bearing steels showed the addn of 0.35-0.45% Cu to diminish the wt. loss in 5% H₂SO₄, 5% HCl and 50% Greater addns. did not increase the protection, which was about 90% in H₂SO₄, 60% in HCl and 80% in H₂O. Addns of 0.36% or more Cu brought about a reduction in tensile-strength loss which was 1/2 that of Cu-free steel. Service tests, in which the plates were located in the roof of a steam railroad tunnel, showed that Cu-bearing steels were about 15-20% more resistant than Cu-free steels. B. E. ROETHLI.

The role of hydrocyanic acid vapors in the corrosion of iron. J. F. G. HICKS, *J. Phys. Chem* 35, 893-904(1931) —Investigations were made to det. the part played by HCN gas in the corrosion of Fe (particularly in the internal corrosion of gas mains), the app. and procedure used being essentially those used previously (C. A. 23, 3199). From graphs of the corrosive effects (mg/sq. cm.) of various gaseous mixts against the time of test (days), it is concluded that: (1) HCN, although not a cause, is the chief contributory factor in the internal corrosion of gas mains, H₂CO₃ and O following in order, and (2) HCN possesses the greatest intensity of effect and is 1st in order of priority of attack. It is shown that in the cases studied HCN does not initiate the corrosion. The mechanism of the "cyanide corrosion" of Fe is similar to other Fe-corrosion processes. Prussian blue is the final product in the corrosion of Fe in the interior of gas mains. The conclusions drawn from the graph are checked electrochemically by the methods in the previous investigation. J. BALOZIAN.

Corrosion "danger" with light metals (aluminum alloys). P. SCHWERRER, *Metallwirtschaft* 9, 158-60(1930) —A comparison of the corrodibility of iron (steel) and of light alloys. The latter are found, generally speaking, much less subject to corrosion. LEOPOLD PESSEL.

The use of aluminum for oil-lease tanks. I. Field tests. LUDWIG SCHMIDT, JOHN M. DEVINE AND C. J. WILHELM *Bur. Mines, Rept of Investigations* 3066, 17 pp (1931). —Al has promising possibilities as a material for the construction of lease tankage because of its resistance to H₂S corrosion. The use of Al steam coils in contact with salt water is not practical. The Al tank plate used did not prove highly resistant to the corrosive action of coned brines. When used in conjunction with steel, extreme care must be taken that no electrolyte is present. ALDEN H. EMBRY.

Condenser-tube corrosion. D. HANSON, *Nickel Bull* 3, 81-5(1930); *Metals & Alloys* 2, Abstracts 64. —A theoretical consideration of corrosion is given, particularly of the mechanism of condenser-tube corrosion, the theory of electrolytic corrosion and the factors influencing corrosion. It is concluded that the soln. of the corrosion problem lies in the use of corrosion-resistant alloys. The possibility of the use of Ni-Cu alloys

is treated. It contains a schematic illustration of a jet test app. used for detg. the comparative corrosion resistance of various condenser tubes. G. G.

A magnetic method for testing boiler tubes. FRANK WEIER AND ARNOLD OTTO. *Mitt. Kaiser-Wilhelm Inst. Eisenforsch. Düsseldorf* 12, 373-87 (1930).—The introduction reviews the history and literature of magnetic testing. Two methods of testing were investigated: (1) the method developed by Sanford and Kourwenhoven, (2) methods developed by the authors. The influence of magnetization field strengths and magnetization in weak and strong fields are discussed. R. RIMBACH.

Atomic hydrogen welding. C. I. MACGUTHRIE. *Baltimore Expr.* 5, No. 5, 15-6 (1930).—General notes on development, process and equipment. E. I. S.

Investigations on the influence of covered welding electrodes on the mechanical properties of the welds. K. BALMGAERTL. *Forschungsarbeiten Geb. Ingenieurw.* No. 336, 33 pp (1930), *Metals & Alloys* 2, Abstracts 70.—The absorption of O and N from the surrounding air during welding reduces the mech. properties of the weld, to exclude or to reduce this influence the welding rods are provided with a kind of covering or are dipped. Tests were made with the following kinds of welding rods: (1) bare rods, (2) dipped rods of the American Welding Process Co., (a) intermediate, (b) normal, (3) asbestos covered rods of the same company, (c) special A, (d) special B, (e) low volt, (f) standard. The rod itself had the following compn. in all cases: C 0.11, Si 0.01, Mn 0.41, P 0.0023, S 0.0077, and the remainder Fe. In the covered rod, asbestos takes the place of a flux. The chem. compn. of the coating is given in detail for every kind of rod used. The 2 light coverings a and b do not materially influence the tensile properties and simply facilitate welding. The asbestos-covered rods covered with a flux, however, show a decided favorable influence as illustrated by phys., chem., metallographic and x ray tests. The least influence was found in the Brinell test, while tensile strength and elongation showed considerable improvement compared with bare electrodes, the same was true also for bending tests, some of the samples could be bent 180° without fracture. X ray tests showed finer-grained structure for covered electrodes. G. G.

Thyratron control equipment for high-speed resistance welding. R. C. GARFINK. *Gen. Elec. Rev.* 33, 511-3 (1930).—In interrupted line welding service requirements of the control device are particularly severe, for this reason application has been made of the Thyratron tube as make-and-break device in control equipment. E. I. S.

Oxy-acetylene welding of copper. W. A. HUBBARD. *Acetylene J.* 31, 250-2 (1930). E. J. C.

Welding corrosion-resistant steels. E. J. TANGEMAN. *Am. Mach.* 73, 45-7 (1930).—Data are included on the phys. properties and compns. of the principal corrosion-resistant alloys. E. I. S.

Arc-welding joints in steel structures. R. W. VAN KIRK. *Am. Metal Market* 38, No. 49 (Mar. 12, 1931). W. H. ROYNTON.

Unusual corrosion problems (POSTER) 14. High-C carbides of the Fe group (FISCHER, BEHR) 6. The galvanic behavior of a Cr-Ni-Fe alloy in sulfite liquors (WESLEY, LAQUE) 23. The emissivity of liquid Fe alloys (NAESER) 3. The system Al-Ag (CREPAR) 2. The crystal structure of the compounds formed in the Sb-Cd system (CHIKASHIGE, YAMAMOTO) 2. Climax Mo deposit of Colorado—with section on metallurgy (BUTLER *et al.*) 8. Quantitative x ray analysis. Cu-Ag and Cu-Zn alloys (TERREY, BARKETT) 3. Thermal diagrams of the systems Ag-Sr and Ag-Ba (WERNKE) 2. A thermodynamic study of the equilibria of the systems. Sb-Bi and Sb-Pb (YAP) 2. Deflocculating (mineral pulp) (U. S. pat. 1,799,277) 13. Electromagnetic separator for separating magnetic from non magnetic material (RUSS. pat. 14,436) 1. Fe oxide and SO₂ (Can. 311,130) 18.

Travail de l'Aluminium et de ses alliages, la fonderie. PARIS. L'Aluminium français. 163 pp.

Travail de l'Aluminium et de ses alliages; la soudure, le rivetage. PARIS. L'Aluminium français. 46 pp.

Metallurgical coke. ALFRED DUFRANSE. *Fr.* 689,738, Feb. 11, 1930. The reactivity of metallurgical coke is reduced and economy of the coke is secured by immersing it in a soln. of lime or cement or by dusting it with powd. lime or cement.

Ores. METALLURGES A-G. *Fr.* 697,583, June 23, 1930. Ores in the form of clumps or mud are freed from their water by mixing them with substances such as coke which allow the water to run off.

Ores of the rare earths. DEUTSCHE GASGLÜHLICHT-AUER G M B H Ger 516,852, May 24, 1927 To facilitate the H_2SO_4 decompn of oxide ores of Zr, Ti and the rare earths, the ores are powd and heated in a rotary furnace with SO_2 and SO_3 . The ores are then pptd with H_2SO_4 and the S recovered as SO_2 and SO_3 as usual. The recovered SO_2 and SO_3 gases may be used for the pre-treatment of further ore in the rotary furnace.

Ore-flotation reagent. IRA H DERBY and ORIN D CUNNINGHAM (to P C Reilly). Can. 310,929, Apr 28, 1931 An ore flotation reagent comprises an oily frothing material and a mono or a di ester or an alkali metal salt of a mono ester of di- or tri-thio-metasilicic acid Cf C A 24, 4752

Apparatus for flotation separation of ores. WM A BUTCHART U S 1,798,451-2, March 31 Structural features

Flotation separation of metallic ores ROBERT W LOYD and BERTCH W MOYE Can 310,001, Apr 7, 1931 An ore pulp is treated by flotation for the extn of Pb products by introducing to the pulp proportionate quantities of NaCl, ZnSO_4 , pine oil and xanthate

Mechanical ore separator. COLORADO IRON WORKS Co Fr 697,428, June 14, 1930

Sluicing apparatus for ore washing. ANTOINE FRANCE Ger 516,860, June 27, 1926

Inclined vibrating concentrating table. MARTIN J LIDF U S 1,799,694, April 7 Structural features

Distributing crushed ore on supports for extraction of metals. HAROLD W ALORICH and WALTER G SCOTT (to Inspiration Copper Co) U S 1,798,715, March 31 App and mech features of operation are described

Preparing materials such as ores for smelting JOHN E GREENAWALT U S 1,799,163, April 7 The material is segregated into 3 or more components of graded sizes, and these are placed in a treating holder in layers in a manner to preserve their respective voids, the component of smallest sized particles constituting the top layer and the components of successively increasing size constituting the succeeding layers and the percentage of voids increasing from top to bottom of the charge App is described

Copper extraction from ore. WM E GREENAWALT U S 1,798,255, March 31. Cu ore is treated to sep a high-grade and a low-grade Cu concentrate. The low-grade concentrate is roasted, the high-grade concentrate is smelted to produce highly coned CuS and the latter is roasted to produce coned. Cu oxide. A portion of the latter is leached with acid to ext the Cu and the soln thus obtained is electrolyzed to deposit Cu and regenerate acid, which is returned to the coned Cu oxide. The roasted low-grade concentrate is leached to ext a portion of the Cu; the resulting Cu soln is treated with a portion of the coned Cu oxide, and the resulting purified Cu soln is electrolyzed with the Cu soln obtained from leaching the coned Cu oxide

Chromium ores. ZAHN & Co BAU CHEMISCHE FABRIKEN G M B H and LUDWIG WICKOF. Ger 516,992, July 14, 1926 Cr ores are disintegrated by heating with an insufficient amt of alkali carbonate and lixiviating the sintered mass. The residue is mixed with 2-3 parts of burned lime or an equiv amt of limestone, 1 part of Fe_2O_3 and alkali carbonate and roasted. An example is given

Treating zinc-bearing concentrates. WM C HOOEY (to New Jersey Zinc Co) U S 1,799,166, April 7 See Can 308,893 (C A 25, 1799)

Treating zinc sulfide concentrates. WALTER O BORCHERT (to New Jersey Zinc Co) U S 1,799,278, April 7 ZnS concentrates contg water-insol Ca compds are treated, as an aq pulp, with H_2SO_4 under such conditions as to convert the Ca compds into minute particles of suspended CaSO_4 without materially attacking the ZnS , and the suspended CaSO_4 particles are sep'd from the ZnS concentrate.

Zinc. ALEXANDER ROITZHEIM and WILHELM REMY U S 1,798,990, March 31. A charge of zinciferous material such as ore and a reducing agent is subjected to external heating, and a small quantity of O is admitted into the charge to increase its temp, thus forming a small quantity of CO_2 . By chem reaction of the CO_2 with some of the Zn vapor, some ZnO in vapor phase is formed, and the ZnO is sep'd from the Zn-metal vapor before the Zn is condensed App is described

Slags. NORSKE AKTIESELSKAB FOR ELEKTROKEMISK INDUSTRI Fr 697,403, June 3, 1930 An app is described for agitating slags to liberate the gases adsorbed or dissolved therein and form bubbles whereby a cellular structure is obtained

Production of metals and their compounds AXEL S BURMAN and IVAR RENNER-FELT. Fr. 698,060, June 24, 1930 A soln of one or more metal compds, or salts,

free from SiO_2 is mixed with a soft substance contg. C, which when heated leaves a residue consisting principally of C. The soft. is evapd. and calcined to form a homogeneous mixt. which is leached with an acid or treated with a gas to sep. one or more of the metals contd. therein. The process may be used to sep. Fe from oxides of other metals such as Al_2O_3 . The calcination step is carried out in an inert or reducing atm. An app. is described.

Spongy metals. HENNING G. FLODRIN. Ger. 522,180, Dec. 25, 1928. Spongy metals are prepd. by briquetting a mixt. of finely ground ore and reducing agent, packing the briquets in a shell or casing, and heating them to a temp. sufficient to reduce the ore without softening the resulting metal. Thus, spongy Fe may be prepd. from a briquetted mixt. of Fe ore, C and a binder at a temp. of 1100° , or spongy Cr from a briquetted mixt. of Cr ore, ferrosilicon and a binder at a temp. of 1200° . Cf. C. A. 24, 330.

Wet-dressing plant for metal foundries. BADISCHE MASCHINENFABRIK & EISENGIESSEIREI VORMALS G. SEHOLD and SEHOLD & NEFF. Ger. 516,793, Jan. 25, 1930. Details of nozzle tubes are given.

Forming sand molds for metal casting. OTTO HARMS. Ger. 516,792, June 20, 1929.

Feeder for ingot molds. BLOOMFIELD H. HOWARD and ERNEST J. TURNER. U. S. 1,800,073, April 7. Feeders are composed of self-supporting combustible material such as wood or paper adapted to burn sufficiently slowly on contact with the molten metal to supply heat to the molten metal within the feeder during the solidification of the ingot, so that the molten metal within the feeder is available to compensate for the shrinkage of the ingot.

Vessel for molding metals. ALGENTINE ELEKTRIKITAET-GES. Ger. 515,943, April 12, 1929.

Die-casting apparatus. NATHAN LESTER (to P. & R. Tool Co.) U. S. reissue 18,014, March 31. Reissue of original Pat. No. 1,673,833 (C. A. 22, 2733).

Casting stereotype plates. SCHNELLPRESSENFABRIK FRANKENTHAL & CIE A-G. Ger. 518,740, Oct. 15, 1925. Details of supplying hot gases are described.

Drawing metals. RUDOLF W. MOLL and HILGO VON BRUCK. Fr. 698,232, July 2, 1930. Metal to be cold-drawn is cleaned with acid, washed in water, treated with a lye of Na_2PO_4 , then dried and drawn, residual mineral oils being used as lubricants.

Open-hearth furnace. NICHAS F. EGGER. U. S. 1,798,871, March 31.

Open-hearth furnace. ROY L. LEVENTHY. U. S. 1,798,611, March 31.

Open-hearth metallurgical furnace. STANLEY H. MCKEE. U. S. 1,798,618, March 31.

Down-blast oil-burning melting furnace suitable for treating ores of iron, lead, mercury, copper, zinc, etc. MATTHEW S. ROGEE (one-half to F. E. Kennedy). U. S. 1,799,643, April 7. Structural features.

Furnace and conveyor suitable for heating small metal articles. FRANK T. COPE (to Elec. Furnace Co.) U. S. 1,799,956, April 7.

Furnace suitable for the heat treatment of metals. FRANK T. COPE and ARTHUR H. VAUGHAN (to Elec. Furnace Co.) U. S. 1,799,957, April 7.

Siemens-Martin furnace. FRIEDRICH SIEMENS A-G. Ger. 516,887, Mar. 9, 1930. The furnace is made from moldable fireproof material.

Furnace for annealing metals. AKT.-GES. BROWN, BOYER & CIE. Fr. 693,497, April 11, 1930.

Continuous-heating furnace suitable for heating metal sheets. SAMUEL E. DIERSCH (to S. Dierscher & Sons). U. S. 1,800,170, April 7. Structural features.

Means for preventing burning of blast-furnace tuyères. ERNEST H. HOLLEWORTH. Ger. 522,179, Oct. 8, 1929. Brit. 331,464 (C. A. 25, 65).

Discharge device for grate shaft furnaces for roasting lump ore. HERMANN GLEICHMANN. Ger. 522,422, Oct. 27, 1929.

Heat treatment of metal articles in fused salt baths. WALTER BECK (to Deutsche Gold & Silber Scheideanstalt vorm. Roessler). U. S. 1,799,945, April 7. Articles such as tools are heated in a bath of fused salts such as NaCl and KCl in the presence of finely divided charcoal at temps. low enough to prevent any chem. reaction between the metal articles and the constituents of the bath. An alk. compd. such as NaOH may also be added.

Use of gas jets for cutting slots in metal plates. CLARENCE J. COBBY (to Kobe, Inc.) U. S. 1,799,612-13-14, April 7. Various details of app. and procedure are described.

Manganese pig iron. SAMUEL B. SHELTON (First and American National Bank

of Duluth, executor). Can 311,152, May 5, 1931. A mixt. of a siliceous ore and basic-open hearth slag, the latter contg Mn, is charged into a blast furnace, and is so proportioned as to give a pig iron contg approx 8% Mn. This pig iron is charged into another blast furnace with Fe ore to give a pig iron contg approx 2% Mn. This latter iron is suitable for making steel.

Reducing oxides of iron. MARIO AMOROSO and SOCIETÀ ANON METALFER. Fr 697,626, June 18, 1930. The reduction of oxides or other compds of Fe is accelerated by submitting the mass, during reduction, to the action of a magnetic field for orientation of the atoms of Fe.

Magnetic iron. I G FARBENTND A G. Fr 698,642, July 8, 1930. Fe with excellent magnetic properties is prepd by heating Fe in pieces (prepd without fusion by pressure or beat or both from finely divided Fe) in a vacuum to a temp below the m p, preferably in an atm of H₂. The finely divided Fe may be obtained from Fe carbonyl.

Steel. GEORGE A. CODE. Can 310,972, May 5, 1931. Is a single-step process of producing steel or its alloys, iron ore is reduced in molten condition while associated with a flux and mucic acid.

Steel. KARL ENGEL. Austrian 121,797, Sept 15, 1930. Corrosion after welding of austenitic Cr-Ni steel is avoided by applying an acid-resistant coating at and about the welded parts.

Steel. M. STEAN A. G. Fr 697,811, June 23, 1930. Valuable steels are made by adding waste or scrap metal contg Cu or Cu and Ni to the fusion bath.

Coloring steel. SCHÖELLER BLECKMANN STAHLWERKE A G (Karl Thomis, inventor). Austrian 121,991, Nov 15, 1930. Steel articles are colored brown by successively steeping in 3 solns. contg (1) oxalic or saccharic acid, c g. a 10% soln. (2) an alkali, alk. earth or NH₄ sulfide or hydro- or poly-sulfide, c g. a 1% soln. of Na₂S, and (3) a manganate or permanganate, c g. a 5% soln. of KMnO₄. The articles may be rinsed after each steeping and should be first polished and degreased. The method is particularly applicable to rustless steel articles. Treatment with H₂S may be substituted for the second steeping.

Thermal treatment of alloys. SOC. ANON. DE COMMUNITY, FOURCHAMBAULT ET DECAEVILLE. Fr. 698,724, Oct. 11, 1929. Alloys contg Ni (or Ni and Co) 6-80, Cr 0-40, Mn 0.3-4, W 0-10, Mo 0-10, V 0-2, Ti 0-0.5, Si 0-3% and Fe the rest in the γ state have their elastic limit increased by tempering between 800° and 1200° and then heating for 2 to 200 hrs. between 400° and 900°.

Aluminum-base alloy. ROBERT S. ARCHER and LOUIS W. KEMPF (to Aluminum Co. of Am.). U. S. 1,799,837, April 7. An alloy which is suitable for engine pistons, etc., comprises Al together with Si 7-15, Mg 0.2-3.0, Ni 0.5-7.0 and Cu 0.3-4.0%. Cf. C. A. 25, 909.

Lead alloys. FELTEN & GUILLAUME CARLSWERKE A. G. Ger. 522,364, Dec. 4, 1926. Pb for sheathing a. c. single-conductor cables is alloyed with more than 5% of Bi and (or) Sb, so as to increase its elec. resistance by at least 25%.

Alloys of magnesium. I G FARBENTND A. G. Fr. 697,747, June 21, 1930. The resistance qualities of alloys of Mg are improved by heating the alloys once or more to a temp of at least 800° before casting. Cf. C. A. 25, 909.

Iron-nickel alloys. ELECTRICAL RESEARCH PRODUCTS, INC. Ger. 516,696, Aug. 25, 1925. A Fe-Ni alloy for magnetic purposes contains 75-83% Ni. A small addn. of Cr or Mo may be present, i. e., up to 7%.

Nickel-chromium-iron alloy. NORMAN B. PILLING (to The International Nickel Co., Inc.). Can. 310,870, Apr. 21, 1931. To an alloy of 1-95% Ni, 1-50% Cr and the rest Fe is added a very small quantity of Ca before pouring to enhance its working properties.

Nickel steel or nickel-iron alloys. THOMAS W. HARDY, JR., and CHARLES B. PARSONS. Can 310,951, May 5, 1931. A sulfide ore contg Ni, Fe and Cu is subjected to selective flotation to provide a Ni-Fe sulfide concentrate. The concentrate is roasted to remove S and the mass agglomerated. The mass is heated to a temp. below its sintering point in a reducing atm. to reduce it to the metallic state, and the metallic product is melted to produce steel or alloy.

Steel alloy. THEODOR CIEPL. Fr 697,801, June 23, 1930. A steel resistant to fire and corrosion contains C up to 1, Ni 25-40, Mo 8, Cu 1-8% and Fe the rest. The Ni may be replaced more or less completely by Co, and the alloy may contain up to 10% of Ta.

Alloy steel resistant to corrosion and to oxidation at high temperatures. PEACV

Soldering aluminum. RAYMOND ALPHONSE FAÏRE *Fr* 698,683, Oct 5, 1929. An alloy of Al, Zn, Sn and Cu with a secondary alloy of Al and Zn, or an alloy of Sn, Zn, Pb and P with a secondary alloy of Sn, Zn and Pb (or Cd) or a single alloy of Zn and Cd (or Sn) is used for soldering Al

10—ORGANIC CHEMISTRY

CLIAS A. ROUTHIER AND CLARENCE J. WEST

The first hundred years of synthetic organic chemistry. F. MASTRANGIOLI *Ann soc pharm chim Sao Paulo* 2, 119-33(1931). —A review M. H. SOULE

Thermal decomposition of organic compounds from the standpoint of free radicals.

I. Saturated hydrocarbons. F. O. RICE *J Am Chem Soc* 53, 1939-72(1931). — When a hydrocarbon decomps we may assume that it dissociates into 2 free radicals which fly apart. Comparison of the strength of a C—H bond (93.3 kg cal) with that of a C—C bond (71.0 kg cal) shows that the dissociation occurs only through rupture of a C—C bond, any breaking of a C—H bond is wholly negligible compared with that of a C—C bond. The probability that the 2 radicals collide again is negligibly small and therefore, free radicals produced in this way can only decompose or react with the surrounding hydrocarbon mols. E.g., a free Me group can take a H atom from the surrounding hydrocarbon and form CH_4 and a hydrocarbon radical. An Et group produced in this way can undergo 2 reactions: the 1st is similar to the Me group and produces C_2H_6 and a hydrocarbon radical; the 2nd is a dissociation into C_2H_5 and H atoms. Radicals higher than Et can decompose in a similar manner into an C_2H_5 hydrocarbon and either a H atom or a free radical. This dissociation of free radicals is possible because, in the process, a single bond becomes a double bond, this process of internal compensation releases 52,800 cal which can contribute to the min activation energy (93,000 cal) necessary to break a C—H bond or to the min activation energy (71,000 cal) necessary to break a C—C bond. The decomposition of paraffin hydrocarbons is represented as a chain type of reaction in which free H atoms or free radicals combine with 1 of the H atoms of the surrounding hydrocarbon mol. The hydrocarbon radical then decomposes into a compd and either a smaller free radical or a H atom. Thus we have a cycle of changes in which certain groups called carriers are regenerated and start a new cycle and this process presumably may go on for a great no. of times. The composition of the products is determined therefore almost exclusively by the chain cycle and is practically independent of the primary decomposition of the hydrocarbon. It is possible to calc. the compn of the products when a hydrocarbon is allowed to decompose to a small extent. In these calcns all primary C—H bonds are considered to have the same strength, all sec. C—H bonds 1200 cal less and all tertiary C—H bonds 4000 cal less than a primary C—H bond. The only arbitrary assumption made in the calcns is the assignment of the relative strengths to the different classes of C—H bonds. The following hydrocarbons are discussed: C_2H_6 , C_3H_8 , iso- C_4H_{10} , C_4H_{10} , C_4H_{10} , iso- C_5H_{12} and CMe_4 .

C. J. WEST

Physical properties of the normal paraffin hydrocarbons, pentane to dodecane. A. F. SHEPARD, A. L. HENNE AND T. MIDGLEY, JR. *J Am Chem Soc* 53, 1948-58 (1931). —The samples were prepd from a special gasoline with a normal paraffin hydrocarbon content of about 70%, this was roughly sep'd into fractions by distn and then treated with ClSO_3H , the progress of the further purification was checked by the total boiling range and the total freezing range, details of the abs temp measurements are given. B ps., m ps., d_4^{25} , d_4^{25} , n_D^{25} , n_D^{25} and n_D^{25} at 20°, 25° and 45.3° each, viscosities at 25°, fluidities at 25°, critical soln temps in amine, and Nekrasov's consts (cf C. A. 23, 3390) are given for C_5H_{12} , C_6H_{14} , C_7H_{16} , C_8H_{18} , C_9H_{20} , $\text{C}_{10}\text{H}_{22}$, $\text{C}_{11}\text{H}_{24}$ and $\text{C}_{12}\text{H}_{26}$.

C. J. WEST

Alternation in properties of long-carbon-chain compounds. T. MALKIN *Nature* 127, 126-7(1931). —The concept of alternation in all long-chain compds is modified so that the zigzag chain is tilted with respect to the terminal planes. Nonalternating series possess vertical chains (hydrocarbons, Me ketones, alcs) alternating series possess tilted chains (mono- and dibasic acids, iodides, nitriles). The concept permits a longer spacing for odd numbered compds than for even ones and greater mol vol. for odd numbered solid fatty acids.

FRANK MARESH

The catalytic addition of gaseous hydrochloric acid to unsaturated hydrocarbons. W. J. PIOTROWSKI AND J. WINKLER. *J. Inst Petroleum Tech* 17, 225-41(1931). —See C. A. 25, 1794

G. G.

Pyrolysis and condensation of hydrocarbons. I. Ethylene. E. BERL AND W. FORST *Z. anorg. Chem.* 44, 193-7(1931).— C_2H_4 is decompd. by recirculating 45 l. of it through a tube furnace. Around 800-900°, 43% of the C_2H_4 is condensed into liquid products giving C_2H_6 , C_4H_{10} , anthracene and their homologs, with traces of cyclopentadiene. Metals and oxide catalysts decrease the yield of liquid products and increase the decompn. of C_2H_4 into its elements. The formation of the liquid product proceeds in 2 steps: the pyrolysis of the mol into free radicals, and the subsequent hydrogenation or condensation of the latter.

GEORGE CALINGAERT

Rate of polymerization. I. Mechanism of polymerization of diethylene hydrocarbons. E. N. GARON *J. Russ. Phys.-Chem. Soc.* 62, 1385-93(1930).—G. discusses the mechanism of polymerization of org. compds. having 2 double bonds and draws the following conclusions: (1) The essence of activation of a hydrocarbon mol. in the process of polymerization consists in the breaking of one double bond and the changing of quadrivalent C atoms at the rupture into trivalent. (2) The theoretical value of the energy of activation is equal to the energy of breaking of one double bond under conditions where chain reactions are absent. (3) Active mols. on reacting with normal ones, give products of polymerization, while they cannot react among themselves. II. Rate of polymerization of isoprene. *Ibid.* 1395-1406.—By using the method of heating samples of isoprene in sealed tubes in a thermostat, in the absence of light, a study was made of rate of polymerization at 100°, 110°, 130° and 150°. In the activation of the mols. of isoprene there appear 3 active forms giving rise to 4 primary products: dipentene, 1,3-dimethyl-3-vinyl-6-cyclohexene, and 2 polymers. An equation is developed for a process which goes on simultaneously in various directions independently of each other. The consts. of rates of reaction of the primary products and their temp. coeffs. are detd. The values of the heat of activation and the reaction consts. for all 3 reactions are calcd. The consts. of the rate of formation of the primary products are in the same order to each other as the heats of formation E and the reaction consts. S .

S. L. MADORSKY

Formation of methanol by the direct oxidation of methane. WILLIAM A. BOVE. *Nature* 127, 431(1931).—Studies on the slow combustion of C_2H_6 by B. and Hill (C. A. 25, 1300) showed that the initial product was not a peroxide but either $EtOH$ or some less oxygenated substance. It has now been found that direct oxidation of CH_4 at 300° and 100 atm. gives rise to $MeOH$ as the initial product. In mixts. of CH_4 and O_2 in the ratio of 9:1 the action was complete in a few min. About 17% of the CH_4 burnt was recovered as $MeOH$, 0.6% as CH_3O and the remainder as CO , CO_2 and H_2O . No H was liberated and not even a trace of peroxide was formed. The $MeOH$ was identified by conversion into $HOC_2H_4CO_2Me$ and $p-NO_2C_6H_4CO_2Me$ and was estd. as the nitrite.

H. W. LEAHY

Action of hydrogen peroxide upon simple carbon compounds. I. Methanol, formaldehyde and formic acid. H. SHIRLEY FAY AND JOHN H. PAYNE. *J. Am. Chem. Soc.* 53, 1973-80(1931).— HCO_2H (I) is oxidized by H_2O_2 to H_2CO_3 and H_2O , 76% of the acid reacting, no H or CO was found. $HCHO$ (II) reacts in 2 ways: $2H + H_2O_2 = 2I + H_2$ (a), $II + H_2O_2 = I + H_2O$ (b), the I is also oxidized, giving $H_2CO_3 + H_2O$ (c). With $1/11$ mol H_2O_2 , no I is oxidized, with $1/2$ mol H_2O_2 , I is only oxidized to a slight extent, with increasing concn. of H_2O_2 , the extent of the occurrence of reactions (c) and (b) regularly increases while that of (a) regularly decreases, the liberation of H is apparently favored by the lower concn. of H_2O_2 . In the oxidation of $MeOH$, the first reaction is $MeOH + H_2O_2 = II + 2H_2O$ (d); the other reactions (oxidation of I and II) also occur, so that the oxidation of $MeOH$ gives I, II, H_2CO_3 and H . The reactions represented by equations (a) and (b) decrease in extent while those represented by (c) and the decompn. of H_2O_2 into O_2 and H_2O increase in extent with increasing concns. of H_2O_2 . Reaction (b) always occurs to a greater extent than reaction (a), again showing that the liberation of H is favored by the lower concns. of H_2O_2 . The results are given in a series of tables. II. The mechanism of the reactions. *Ibid.* 1980-4.—In the oxidation of AcH the yields of the 2 chief reaction products, $AcOH$ and II_2CO_3 , increase as increased quantities of H_2O_2 are used. Since separately conducted runs with $AcOH$ and H_2O_2 gave HCO_2H and H_2CO_3 without the liberation of H or of CH_4 , it follows that the II arises from the oxidation of AcH to $AcOH$ ($2AcOH + H_2O_2 = 2AcOH + H_2$). While the max. yields of II were obtained with equimol. concns. of AcH and H_2O_2 , another striking parallel to the $HCHO$ reaction is found in the fact that the vols. of CH_4 liberated according to the proposed equation, $2AcH + H_2O_2 = AcOH + HCO_2H + CH_4$, decrease with increasing concns. of H_2O_2 . The formation of $AcOH$ and HCO_2H with the concurrent liberation of both H and CH_4 appears to be in conformity with the equations for the reactions which

were predicted in extending the proposed reaction mechanism to AcH . The term *perhydrolysis* is defined as a double decomposition reaction involving H_2O_2 in precisely the same manner that hydrolysis is a double decomposition reaction involving H_2O . The reduction of $\text{H}_2\text{C}_2\text{O}_4$ by H_2O_2 with the liberation of O may likewise involve perhydrolysis, forming peroxanthonic acid, which, in turn, yields HCO_2H through the liberation of O . All reactions of H_2O_2 do not involve perhydrolysis.

C. J. WEST

Isomeric Isoprene sulfone. H. LEIST, EIGENBERGER, *J. prakt. Chem.* 120, 312-20 (1931), cf. C. A. 25, 276. The action of 0.5 *N* aq. or alc. KOH upon isoprene sulfone A (I) in the light of a quartz light for 20 min. gives isoprene sulfone B (II), m. 70° , which is less sol. in H_2O than its isomer. In the absence of KOH light has no action, the 2 isomers form a mixt. m. 30° . It does not add to H in the light or at 100° . Heating with *N* 1.0 *N* KOH for 2 hrs. causes about 50% sapon, giving an amorphous, brown substance. Oxidation with *N* KMnO_4 in K_2CO_3 gave no definite cryst. product. I and K in dry H_2O give the K salt very hygroscopic, of 3,6 dimethyl-2,4 octadiene-1,8 dithiolonic acid, oily, very unstable and reactive, the most pronounced property of which is its tendency to polymerize, the products being an H_2O insol., AcOH sol. oil and an amorphous compd., insol. in the ordinary solvents. The acid gives an unstable H salt complex, the oily polymer also gives a sol. linimide, but this was not investigated because of its instability. The Ca salt and Zn salt, m. $70-5^\circ$ were also prepd. The K salt and MeI heated 1 hr., give dimethyl isoprene disulfone h_{11-11} $130-40^\circ$, which polymerizes in part during the distn. Oxidation of the K salt gives H_2SO_4 and acetylacetone. II gives a sulfonic acid which yields the same oxidation products. I and K in moist H_2O give $\text{Me}_2\text{C}(\text{CH}_2\text{CH}_2\text{SO}_2\text{H})$.

C. J. WEST

α -Chloroethane- α -sulfonyl chloride. ERNST MCLIFF and HERMANN RAUBENUSCH, *Ber.* 64B, 91-101 (1931).—Chlorination of β -trithioacetaldehyde (I) in aq. suspension had given an oil whose reactivity toward NH_3 , PhNH_2 , NaOH , and PhNHNH_2 indicated that it was an ethane-sulfonyl chloride, a conclusion confirmed by analysis (C. A. 21, 2472). A more thorough study has shown that it h_{11} 70° and is not identical with the known $\text{ClCH}_2\text{CH}_2\text{SO}_2\text{Cl}$. It can therefore be only α -chloroethane- α -sulfonyl chloride (II). It is also obtained in glacial AcOH , but in dry CCl_4 there is formed only a dark yellow oil, $\text{C}_2\text{H}_3\text{Cl}_2\text{S}$ (III), which is very probably a heptachloroethyl sulfide. In H_2O , along with II are formed H_2SO_4 and AcOH and the yield of II is only 50%. It is also formed with HOCl but the formation of H_2SO_4 and AcOH is greatly favored in this reaction, the preformed II is also much more readily oxidized by free HOCl than by Cl in H_2O . II readily dissolves in the common org. solvents but boiling H_2O converts it into the acid (IV). It is sol. in NaOH and in NaOH and with NaOH gives the amide (V). With H_2O in sealed tubes it gives only IV. Boiling H_2O does not attack II while hot anhyd. MeOH gives IV, possibly through the Me ester (VI) as an intermediate product which is saponified by traces of H_2O formed by the action of the liberated HCl on the MeOH . EtOH in sealed tubes at 130° likewise gives IV instead of the ester. With NaOMe instead of MeOH , VI can be obtained in almost pure form, together with considerable amts. of the Na salt of IV, probably formed according to the equation $\text{MeCHClSO}_2\text{Me} + \text{MeONa} \rightarrow \text{MeCHClSO}_2\text{Na} + \text{Me}_2\text{O}$. With PCl_5 at 200° in sealed tubes, II gives MeCHCl_2 , at room temp. or on heating in an open test tube no reaction occurs. Like other sulfonyl chlorides II is formed from the Na salt of IV with PCl_5 and is reduced to the mercaptan by nascent H , but with Zn dust it gives EtSO_3K (VII), instead of a sulfinate when the mixt. is neutralized with K_2CO_3 at the end of the reaction. Traces of H_2O are essential for the reaction, which takes place in alc. and Me_2CO but not in abs. anhyd. media like CHCl_3 and C_6H_6 , in abs. H_2O it sets in immediately when a drop of H_2O is added. III, h_{11} $35-8^\circ$, V, m. 66° , Phenylhydrazide IV, thick, gray brown liquid which retains H_2O after long standing over H_2SO_4 , VI (impure, C. 21.93, H. 4.23, Cl. 23.73%), h_{11} $91-3^\circ$, VII, hitherto described as crystals with 2 mols. H_2O , was obtained in anhyd. needles.

C. A. R.

Ethylene nitriles. The 2-methyl-3-pentene-5-nitriles, P. BRIVANETS and H. MINETTI, *Bull. sci. acad. roy. Belg.* [5], 16, 1116-25 (1930).—Henry (Chem. Zentr., 1898, 11, 662) described 2-methyl-3-pentene-5-nitrile, $\text{Me}_2\text{CHCH}(\text{CN})\text{CH}=\text{CH}_2$ (I), as a liquid b. $154-5^\circ$, but his material was a mixt. of the 2 geometric isomers and 3-methyl-3-pentene-5-nitrile because he started with ordinary fusel oil. Iso AmOH (II), obtained in 4.4 kg. yield by hydrolysis of 15.0 kg. ($\text{Me}_2\text{CHCH}(\text{CH}_2\text{OSO}_2\text{R})\text{CH}=\text{CH}_2$, 211.0 (prepd. from fusel oil by the method of Marekwall) with H_2SO_4 , h_{11} $131.5-2^\circ$ (polariscope measurements indicated 98% active alc.), d_{40}^{20} 0.80788. $\text{Me}_2\text{CHCH}(\text{CH}_2\text{OH})\text{CH}=\text{CH}_2$ (III), obtained in 1550 g. yield by oxidizing 4.4 kg. II in 130 g. lots with 110 g. $\text{Na}_2\text{Cr}_2\text{O}_7$,

220 g H_2O and 140 g 66° BE H_2SO_4 (yielding about 35% III and 35% II returned for oxidation), $b_{\text{m}} 91.9-2.1^\circ$, $d_4^{20} 0.79770$, $n_D^{20} 1.38710$, 1.38730, 1.39121, 1.39831 for α , ν , β and γ at 20° . $\text{Me}_2\text{CHCH}_2\text{CH}(\text{OH})\text{CN}$ (IV), obtained in 2150 g yield from 1870 g III, $b_{\text{m}} 107^\circ$, $d_4^{20} 0.91852$, $n_D^{20} 1.42317$, 1.42514, 1.43018 for α , ν and β at 20° . I was obtained from IV by dehydration with P_2O_5 in CHCl_3 but in better yield (64%) and more easily purified without the use of a solvent. Repeated fractionation of about 500 g I yielded chiefly 2 fractions: (1) $b_{\text{m}} 137.5-8^\circ$, $d_4^{20} 0.89781$, $n_D^{20} 1.42180$, 1.42493, 1.43246, 1.43891 for α , ν , β and γ at 20° , probably the *cis* modification of I, and (2) $b_{\text{m}} 160-0.5^\circ$, $d_4^{20} 0.82131$, $n_D^{20} 1.43290$, 1.43620, 1.44390, 1.45040 for α , ν , β and γ at 20° , probably impure *trans* I since both correspond on analysis to $\text{C}_{11}\text{H}_{19}$, and appreciable quantities of a higher boiling material (VI) where the exaltation of the mol refraction decreases with the h. p., indicating the presence of 2 methyl 2 pentene 5-nitrile (V), formed by isomerization of I. The I absorption method of Kandiah and Linstead (cf C A 24, 76) and a better Br absorption method to indicate the relative amts of type I and V nitriles confirmed the structure assigned to V. The Br absorption method applied to *cis* I indicates 0.75% (within the accuracy of the method) of V in it, 21% in impure *trans* I and 36% in a fraction $b_{\text{m}} 164.5^\circ$. Fractionation of VI under reduced pressure was not successful in sepg V, and fractional crystn at -89° to -110° did not appear promising (from the difference in the rate of Br absorption), though the lower melting material is somewhat higher in V. To a soln of 40 g VI ($b_{\text{m}} 160-0.5^\circ$) in 60 cc CHCl_3 in an ice bath was added 27 g Br (about 100% excess), the excess Br removed with bisulfite soln and the CHCl_3 soln dried and distd, the product, $b_{\text{m}} 44.5-0^\circ$, $b_{\text{m}} 156-7^\circ$, $d_4^{20} 0.82131$, $n_D^{20} 1.43075$, 1.43419, 1.44190, 1.45190 for α , ν , β and γ at 20° , is probably pure *trans* I. It undergoes slight decompn when distd at atm pressure and its h. p. and d rise. This repeated distn is not responsible for the formation of V, but the dehydration with P_2O_5 produces it as shown by a sep expt: *cis*- and *trans* I add HCl easily to form $\text{Me}_2\text{CHCH}_2\text{CH}(\text{OH})\text{CN}$ (VII), $b_{\text{m}} 91.4$ (probably error for 90.4° —Abstractor) -1.4° , $d_4^{20} 1.0189$, $n_D^{20} 1.43775$. Treatment of 60 g VII with quinoline gave 31 g crude unsatd. nitrile, most of which $b_{\text{m}} 159-61^\circ$, the Br absorption indicated 17.6-21.6% V and probably the remainder was *trans* I as the h. p. did not indicate any *cis* I. Re fluxing for 2 hrs 30 g impure *trans* I ($b_{\text{m}} 160-0.5^\circ$) and by Br absorption contg 21% V with 1 g phenol and 0.1 g Na yields a product, $b_{\text{m}} 164-6^\circ$, $d_4^{20} 0.82873$, $n_D^{20} 1.43634$ and by Br absorption contg 43% V. Refluxing 28 g *cis* I with PhONa under the same conditions gave a fraction $b_{\text{m}} 161.6^\circ$, $d_4^{20} 0.84082$, $n_D^{20} 1.43400$ and by Br absorption contg 51.8% of V, while a fraction $b_{\text{m}} 166-8^\circ$ contained 67.7% V. Ultra violet light does not bring about isomerization as does the PhONa : *trans* I was converted by treatment with H_2SO_4 for 14 days, into the corresponding amide, m $87.6-8.2^\circ$. *cis* I required 60 days with H_2SO_4 to produce the corresponding amide, m 80° . Equal amts of the 2 amides m $55-61^\circ$.

ODEN E. SHEPPARD

Production of s-trichloroaniline and tetrachloroquinone from aniline in hydrochloric acid solution. J EADÉLYR *Chem Rundschau Mitteleuropa u. Balkan* 6, No 14, 83-4 (1929), cf C A 23, 4890.—Aniline is dissolved in concd HCl and chlorinated at $35-40^\circ$. s-Trichloroaniline is formed without any colored or resinous secondary products. Tetrachloroquinone is formed if aniline in HCl is treated with a mixt of Cl and O .

S. S. DE FINAY

The preparation of trialkylbismuthines. W. C. DAVIES, I. NORVICK and W. J. JONES *Bull soc chim* [4], 49, 187-94 (1931).—Tripropylbismuthine— Pr_3MgBr was prepd by the addn. of 61 cc. of PrBr in 100 cc. Et_2O to 16.2 g of Mg in 100 cc Et_2O in a specially arranged Grignard app. permitting the maintenance of a H_2 atm., and also the withdrawal of the sepd. layers into a separatory funnel provided with CO_2 intake and outlet tubes. To the cooled Grignard reagent, under H_2 , was added dropwise 30 g BiCl_3 in 250 cc Et_2O . The mixt. was heated for 30 min., then cooled and after decompn. with aq. NH_4Cl the ether layer was transferred to the CO_2 -filled separatory funnel. After drying over Na_2SO_4 and evapn. of the Et_2O , both operations being carried out in a CO_2 atm., the residue was fractionated *in vacuo*, the main fraction consisting of 16.5 g of clear liquid, $b_{\text{m}} 56-7^\circ$, $d_4^{20} 1.621$. By operating in a closed box provided with flexible sleeves for the arms of the operator, and a glass observation window, the regulation of this fraction was carried out in a CO_2 atm., tubes for C , H and Br analysis being filled and sealed in this way. By similar procedure, triethylbismuthine, $b_{\text{m}} 117.1-4.5^\circ$, and triamylbismuthine, $b_{\text{m}} 157-8^\circ$, $d_4^{20} 1.381$, were prepd. Some hexane, 3 g, $b_{\text{m}} 68-70^\circ$, and 10 g of decane, $b_{\text{m}} 160^\circ$, were formed by

the synthetic action of the Mg. Though the bismuthines obtained as main fractions of distn were remarkably pure, bismuthine halides were produced when the Grignard reagent was not in 2-fold excess. The trialkylbismuthines are clear liquids which oxidize readily and are thermally unstable. Their b p's rise progressively in the series but their ds diminish. Tables of b p, b p reduced to 50 mm and ds, are given, together with a résumé of the comparative chem. and phys. properties of the trialkyl derivs. of the elements of the 5th group. C. R. ADDINALL

New compounds of titanium. H. A. GARDNER AND E. BIELOUSS. *Am Paint & Varnish Mfrs' Assoc., Circ No 366, 327-37* (1930).— $Ti(OEt)_4$ prepd. by dropwise addn of 132.6 g $TiCl_4$ to a cold soln of NaOEt with subsequent refluxing, contained 31.7% TiO_2 , calcd 31.5%. The product was sol in alc, C_6H_6 , toluene and mineral spirits. $TiCl(OAm)_3$, HCl and $Ti(OAm)_4$ were prepd by dropwise addn of 19 g $TiCl_4$ to 40 g pentasol (mixed amyl alcs contg 26% secondary alcs). The product was dark and viscous. Distn (30 mm) below 100° yielded 12 g alc., from 100° to 240° it yielded 5 g $Ti(OAm)_4$, which hydrolyzed rapidly in the air. Ti compds of glycol and glycerol of varying compns were prepd by dropwise addn of $TiCl_4$ to an excess of the corresponding polyhydric alc at temps from 80° to 145° . The products were sol in water and alc but not in toluene. Films of $Si(OEt)_4$, hydrolyzed as directed by King (*C. A.* 24, 2313) and of $Ti(OEt)_4$, hydrolyzed *in situ*, applied to various surfaces were unsatisfactory as they lacked adhesion and continuity. $TiCl_4$ converts tung oil to a gelatinous mass. A white-pigmented lacquer was formed when Ti glycolate in alc. (1.4) was added to nitrocellulose in AcOBu, the Ti compd hydrolyzing with the formation of TiO_2 . $Ti(OEt)_4$ gave unsatisfactory results as an antiknock agent in gasoline because of pptn of TiO_2 on the walls of the passages between the carburetor and the cylinders. G. G. SWARD

State of formaldehyde in aqueous solutions. FREDERIC WALKER. *J. Phys. Chem.* 35, 1104-13 (1931).—The object of W's paper was to correlate the scattered data on the subject of the state of CH_2O in H_2O . In doing this, W has found that the theories and data of Auerbach (A. and Barschall, *Arb. Kauts. Ges. A.* 22, 584-629) enable one to explain quantitatively the vapor pressure of CH_2O in H_2O as recently detd by Ledbury and Blair (*C. A.* 19, 1214). In addn W has shown (1) that such solns follow Henry's law, in that the partial pressure of the CH_2O vapor is proportional to the concn of solute in the form of $CH_2(OH)_2$, and (2) that the gas in equil. with aq solns is probably the unhydrated monomol CH_2O . LOUISE KELLEY

Action of ozone upon aldehydes. F. G. FISCHER, H. DELL AND J. L. VOLZ. *Ann.* 486, 80-94 (1931), cf *C. A.* 24, 2106.—The action of O_3 on aldehydes gives a mixt of the corresponding acid and peracid; it may be assumed that the aldehyde adds 1 mol O_3 and that the $RCHO$, O_3 then reacts with 1 mol $RCHO$ to give RCO_3H and RCO_2H ; with 2 $RCHO$ to give 3 RCO_2H or with O_3 to give RCO_3H and O_2 . The best yields of peracids are obtained by ozonization of small amts of aldehyde in dil soln at -15° to -20° . Thus $PrCHO$ (8 g) in 400 cc. CCl_4 gives 71.8% $PrCO_3H$, b_p 26-30°, 150-BuCO₃H, b_p 31-2° (92.7% yield), $BzCO_3H$, m 39-40° (2.65 g yield from 8 g BzH). The yield of peracid is affected by the diln of the aldehyde, by the temp and by the solvent, the yield decreases with increasing concn, a higher yield of peracid is obtained at 10° than at -65° . C. J. WEST

Reactions of certain carboxylic acids. LAD EKKERT. *Pharm. Zentralhalle* 72, 228-9 (1931), cf *C. A.* 25, 2389.—The behavior of HCO_2H , AcOH, propionic, butyric, isobutyric, valeric, capric, caprylic, palmitic, stearic, oleic and ricinoleic acids has been studied with respect to their behavior toward alc. solns (1%) of CH_2O , $PhCH_2CHO$, $BzCHO$, furfural, anisaldehyde, salicylaldehyde, vanillin, $PhCH_2CHCHO$, piperonal and sucrose when treated with 1 cc coned H_2SO_4 . A long series of color changes is reported. W. O. E.

Lignoceric acid from the seeds of *Adenanthera pavonina*. I. P. RAMASWAMI AYYAR. *Proc. 15th Indian Sci. Cong.* 1928, 161.—The above seeds yield 14% of a fixed oil equiv. to 13% of mixed acids. From the latter pure lignoceric acid, $C_{24}H_{48}O_2$, m $80-1^\circ$, has been prepd. with a yield of 1.5% on the seeds. The acid on bromination by Volhard's method yields a Br acid, $C_{24}H_{47}BrO_2$, m $70-1^\circ$. The lignoceric acid thus appears identical with the acid previously obtained from other sources. E. J. C.

Synthesis of straight-chain unsaturated acids. R. BHATTACHARYA AND J. L. SIMONSEN. *Proc. 15th Indian Sci. Cong.* 1928, 153.—It has been recently suggested that isocroic acid is an inseparable mixt of 2 unsatd acids which may be synthesized from (a) octene, and 14-iodotetradecanoic acid, and (b) hendecen and 10-bromo-undecylic acid. As a preliminary to their prepn the synthesis of homologous acids has been tried. 12-Iodododecanoic acid was prepd. as follows and condensed with the

Na compd of decine to yield behenic acid castor oil \rightarrow undecylenic acid \rightarrow 10-bromoundecylic Me ester \rightarrow cyanoundecylic Me ester \rightarrow decamethylene dicarboxylic Me ester, K salt of the half hydrolyzed ester \rightarrow 12 hydroxy lauric acid \rightarrow 12 iodolauric acid Decine was prepd by the action of $\text{C}_6\text{H}_5\text{Na}$ on octyl iodide. No details are given. E. J. C.

Constitution of hydroxyoleic acid from oil of ergot. HERMANN MATTHIES AND OTTO H. KIRSCHNER. *Arch. Pharm.* 269, 88-101(1931)—Hydroxyoleic acid, isolated for the 1st time by M. and Schutz (C. A. 21, 4017) from oil of ergot, is available in large quantities from this source. The ozonide of this acid is split by water on heating into azelaic aldehydoacid and β hydroxypelargonic aldehyde, which are subsequently converted into azelaic acid, β hydroxypelargonic acid, and a 2nd undetd liquid acid. The HO acid of oil of ergot was recognized as 9-octadecen-12-olonic acid, $\text{Me}(\text{CH}_2)_8\text{CH}(\text{OH})\text{CH}_2\text{CH}(\text{CH}_2)_7\text{COOH}$, identical with ricinoleic acid. Both acids possess like chem. and phys. consts., yielding like derivs. on oxidation with KMnO_4 according to Hazura and splitting of their ozonides. The β hydroxypelargonic acid resulting from the scission of hydroxyoleic acid is unexpectedly l-rotatory, while that obtained from ricinoleic acid has been reported as d-rotatory (cf. following abstr.). W. O. E.

Scission products of ricinoleic acid ozonide. HERMANN MATTHIES AND OTTO H. KIRSCHNER. *Arch. Pharm.* 269, 101-4(1931)—The ozonides and peroxonides of ricinoleic acid and its Me ester are all liquid. Furthermore, the azelaic half-aldehyde yields with NaHSO_3 no cryst. solid, thus preventing a sepn. of this aldehyde from hydroxypelargonic aldehyde, as carried out by Haller and Brochet (C. A. 4, 2091). Scission of ricinoleic acid ozonide yielded in addn. to β hydroxypelargonic acid a liquid acid of insufficient purity for purposes of identification. Contrary to the findings of Haller and Brochet the β hydroxypelargonic acid from the ricinoleic acid ozonide is not d-, but l-rotatory. W. O. E.

Hydrogenation of linolenic acid. H. VANDER VEEN. *Chem. Umschau Fabr., Oele, Wacke u. Harze* 38, 89-96(1931)—V. hydrogenated the Me ester of linolenic acid, prepd by Rollet's method (Z. physiol. Chem. 62, 411(1919)) from the solid hexabromostearic acid of linseed oil. He hydrogenated for 20 min. (for 1 mol. H) at 180° in a Normann beaker with 1.0-1.5% Ni, using a Ni kieselguhr catalyst with 20% Ni (reduced at 500°), hardening until 1 and 2 mols. H had been absorbed. The resulting products may contain (1) unchanged linolenic acid, (2) octadecadienic acids, (3) octadecenic acids, (4) stearic acid. The aints. of these could be calcd. from the 3 known quantities. I nos. (Wijs), percentage of satd. acids (Bertram method), and the Kaulmann thiocyanate nos. The latter were found to be too low and too irregular for a quant. calcn.; only qual. results were obtained by ozonizing the unsatd. acids at the end of the hydrogenation and by analyzing the ozonides formed. The methods of ozonation and the derivation of the final conclusions are clearly outlined. **Conclusions**—The original linolenic ester showed a double bond at the 9,10-, 12,13- and 15,16-C atoms. In the hydrogenation with 1 mol. H the double bond 12,13- is first satd., forming the 9,10,15,16- linoleic acid, then by a partial shifting of these double bonds, a quantity of the isomeric 10,11,14,15- linoleic acid is formed. In the hydrogenation with 2 mols. H the newly formed 10,11,14,15- linoleic acid satd. its 14,15-bond, forming a 10,11- elaidic acid, the 9,10,15,16- linoleic acid satd. its 15,16-bond and should thereby form the normal 9,10-oleic acid, but instead the 9,10-bond is shifted to 8,9 and possibly 8,10 to 10,11, forming isomeric oleic acids. Extended comments are made on the H. Dutch-Vidyarthi paper (C. A. 23, 2422) and the Baner-Erdmann paper (C. A. 24, 5721). P. ESCHER.

Preparation of aliphatic amides. JAMES A. MITCHELL AND E. EMMET REID. *J. Am. Chem. Soc.* 53, 1879-83(1931)—Amides may be prepd by passing NH_3 gas through aliphatic acids kept at a suitable temp. in such a way that the H_2O formed is continually removed, the equl. is displaced far to the amide side. The reaction velocity is appreciably slower with the higher acids, lauric acid, heated at 185° for 16 hrs., gave only a small fraction of the amide, no amide was obtained with palmitic or stearic acid at 125° or 190° . ZnCl_2 did not catalyze the reaction with the lower acids. AcOH , 8 hrs. at 190° , gives 96% of the amide, EtCO_2H , 5 hrs. at 190° , 95%; PrCO_2H , 7 hrs. at 185° , 88%, valeric, 15 hrs. at 180° , 82%, caproic, 7 hrs. at 160° , 75%, heptonic, 7 hrs. at 160° and 4 hrs. at 190° , 75%, caprylic, 11 hrs. at 180° , 80%. Dimethyl-amides may be obtained in a similar way, using Me_2NH , the following yields are reported: HCO_2H , 3 hrs. at 95° , 73%, b. $130-5^\circ$, AcOH , 3 hrs. at 150° , 84%, b. $165-75^\circ$; EtCO_2H , 3 hrs. at 155° , 78%, b. $165-78^\circ$, PrCO_2H , 2.5 hrs. at 155° , 84%, b. $180-94^\circ$; valeric, 3 hrs. at 165° , 87%, b. $205-15^\circ$, caproic, 3 hrs. at 155° , 88%, b. $220-30^\circ$;

heptole, 3.6 hrs. at 160°, 81%, n_D^{20} 1.65–75°. Cor in ps of the amides are given

C. J. WEST

Mechanism of organic reactions. III. The nature of the mechanism of migration of the acyl radical. HAROLD HERRERT AND MARGARET E. GREEN *Can J Research* 4, 251–63 (1931), cf. *C. A.* 24, 4266.—The mechanism first proposed by Emil Fischer to explain the migration of acyl radicals in the case of glycerols of fatty acids receives strong confirmatory support in the isolation by H and G of a product analogous to the intermediate dioxolane deriv. postulated by him. From the theoretical consideration it seemed highly probable that with any glycol monoester contg. a C O group of very pronounced negative polarity, the tendency to undergo cyclization would be so pronounced as to permit of the isolation of the corresponding dioxolane ring. It is shown that all attempts to synthesize $\text{HOCH}_2\text{CH}_2\text{O}_2\text{CCl}_3$ gave, instead, the ring isomer, viz., 2-hydroxy-2-trichloromethyl-1,3-dioxolane. The latter is quite stable at temps below 60°, but at 100–110° decomposes into ethylene carbonate, CH_2Cl_2 , and a certain amt of higher-boiling, unidentified products. In presence of a trace of pyridine the decomposition into CH_2Cl_2 and ethylene carbonate takes place smoothly and apparently quantitatively. Similarly, when ethylene oxide is treated with $\text{ClCH}_2\text{CO}_2\text{H}$, 2-hydroxy-2-dichloromethyl-1,3-dioxolane is produced. On the other hand, AcOH and $\text{CH}_2\text{ClCO}_2\text{H}$ react to give the normal half-esters. 1-pichlorohydrin and glycerol on treatment with $\text{CCl}_3\text{CO}_2\text{H}$ give the corresponding ring compounds, viz., 2-hydroxy-2-trichloromethyl-4-chloromethyl-1,3-dioxolane and 2-hydroxy-2-trichloromethyl-4-hydroxymethyl-1,3-dioxolane, resp. The last-named compd thus represents the cyclic deriv. postulated by Fischer as the intermediate product formed in acyl migration occurring in the case of glycerol deriva. $\text{CH}_2(\text{CH}_2\text{OH})_2$, treated with CCl_3COCl gives 2-hydroxy-2-trichloromethyl-*m*-dioxane. They are all mobile, pleasant-smelling, transparent liquids which decompose when heated at ordinary pressure. The reaction of AcOH , $\text{CH}_2\text{ClCO}_2\text{H}$ and $\text{CCl}_3\text{CO}_2\text{H}$ on 1,2-anhydromannopyranose and 1,2-triacetylmannopyranose is under investigation. Preliminary work appears to indicate the formation of the corresponding 1,3-dioxolane deriv. of mannose in each case. The conclusion is drawn that all org. monoesters of multivalent alcs., carbohydrates, polysaccharides, α -aminophenols, etc., contg. a free OH group, the H atom of which is spatially in close proximity to the C O group of the acyl radical, tend to pass over into a ring isomer. Theoretically, at least, there must exist in all of these cases an equil. between the open chain and its corresponding cyclic isomer. Investigations of the structure of the monoacyl deriva. of glycerol, of carbohydrates, polysaccharides, aminophenols, etc., should permit of the isolation, in many cases, of the intermediate dioxolane or other ring concerned in the migration of the acyl radical. The existence is postulated of a definite equil. between the open chain ester and the corresponding ring isomer in all monoesters of poly-alcs. in which spatial proximity of a OH to the C O group exists, and the prediction is made of the probable isolation of many new cyclic isomers under investigation when appropriate consideration is given to their phys. and chem. properties, such as stability, behavior toward acids, alkalis, etc.

J. W. SMILEY

Physicochemical investigation of amino acids. IV. GAKUJI TAKAHASHI, TOKUZO YAGINUMA AND KENTARO HAYAKAWA *Proc Imp Acad Tokyo* 7, 57–60 (1931), cf. *C. A.* 24, 4266.—*l*- and *d*-I-leucine have practically the same crystallographic constants: the ΔV value for the 2 are $a \ b \ c = 15139 \ 110025, \alpha \ 15331, \beta \ 15361, \gamma \ 15514, V \ 23^\circ 27'$. The values for *d*- and *l*-norleucine are $a \ b \ c = 13921.1 \ 24662, \alpha \ 15071, \beta \ 15101, \gamma \ 15100, V \ 18^\circ 0'$. Histamine picrate, monoclinic, $\beta \ 100^\circ 30', a \ b \ c = 12216 \ 117226$, the values for α, β and γ for $\lambda \ 6659, 6919$ and 4330 are $16851, 16022, 17122, 17009, 17135, 17308, 17532, 17535, 17508$. Histamine forms a mono HCl salt.

C. J. WEST

The behavior of polypeptides, containing lysine with substitution in α - and ϵ -position, toward N alkali, crepsin and trypsin. EMIL ANDERHAGEN AND FRIEDRICH SCHWITZER *Fermentforschung* 12, 330–75 (1931).—The $\epsilon\text{-NH}_2$ group of the lysine component of proteins is generally believed to be unsubstituted, although no convincing proof has yet been offered. This paper deals with the enzymic cleavage of lysine peptides in which the α - or the $\epsilon\text{-NH}_2$, or both, occur in acid amide linkage. The α lysine peptides were obtained by the usual peptide synthesis, starting with ϵ -benzoyllysine in which the $\epsilon\text{-NH}_2$ was already occupied by H₂ and thus excluded from the coupling reaction. Similarly, the ϵ peptides were made from α -methyllysine in which only the $\epsilon\text{-NH}_2$ was free to react. The α, ϵ diaminoacyllysines were obtained directly from unsubstituted lysine. ϵ -Benzoyllysine (I), which served as the starting material, was prepd from benzoylpiperidine by the v. Bmum method

(*C A* 3, 1273) $1 \rightarrow \epsilon$ -N-benzoyl α -(dl α -bromopropionyl)-dl-lysine (II), m 129-30°, $\rightarrow \epsilon$ -N-benzoyl α -dl-alanyl-dl-lysine (III), m 240°, $\rightarrow \epsilon$ -N-benzoyl α -(N-benzoyl-dl-alanyl)-dl-lysine, m 145-6°. $1 \rightarrow \epsilon$ -N-benzoyl α -(dl α -bromoisocaproyl)-dl-lysine (IV), m 148-50°, $\rightarrow \epsilon$ -N-benzoyl α -dl-leucyl-dl-lysine, m 232-3°, $\rightarrow \epsilon$ -N-benzoyl- α -(N-benzoyl-dl-leucyl)-dl-lysine, m 155-6°. IV + $\text{PhCH}_2\text{NH}_2 \rightarrow \epsilon$ -N-benzoyl α -(N-benzoyl-dl-leucyl)-dl-lysine, m 190-1°. $1 \rightarrow \epsilon$ -N-benzoyl α -(dl α -bromocaproyl)-dl-lysine, m 140°, $\rightarrow \epsilon$ -N-benzoyl α -dl- α -leucyl-dl-lysine, m 241-2°. III $\rightarrow \epsilon$ -N-benzoyl α -(dl α -bromoisocaproyl-dl-alanyl)-dl-lysine (V), m 154-5°, $\rightarrow \epsilon$ -N-benzoyl- α -(dl-leucyl-dl-alanyl)-dl-lysine (V) decomps 90°. $1 \rightarrow \alpha$, α -dibenzoyl-dl-lysine, m 146-7°, \rightarrow I, m 252°, \rightarrow II, hence this method is unsuitable for the prepn. of α -benzoyllysine. dl- α -Benzoylaminoo- α -laminocaproic acid (v Braun) + $\text{MeNH}_2 \rightarrow \epsilon$ -N-benzoyl α -methyl-dl-lysine, m 233-4° $\rightarrow \alpha$ -N-methyl-dl-lysine HCl salt (VI), m 244-5° (III salt m 239-41° (decompn)) $\rightarrow \epsilon$ -(dl α -bromopropionyl)- α -N-methyl-dl-lysine \rightarrow ϵ -dl-alanyl- α -N-methyl-dl-lysine (VII), m 115°. VI $\rightarrow \epsilon$ -(dl α -bromoisocaproyl)- α -N-methyl-dl-lysine, m 119-21°. $\rightarrow \epsilon$ -dl-leucyl- α -N-methyl-dl-lysine (VIII) decomps 120°. dl-Lysine mono HCl salt (IX), m 235-6°, is more easily handled than the di HCl salt because of its lower soly and greater readiness of crystn. IX $\rightarrow \alpha$ - ϵ -di-(dl-bromopropionyl)-dl-lysine $\rightarrow \alpha$ - ϵ -di-(dl-alanyl)-dl-lysine (X) IX $\rightarrow \alpha$ - ϵ -di-(dl-bromoisocaproyl)-dl-lysine, m 93-4°, $\rightarrow \alpha$ - ϵ -di-(dl-leucyl)-dl-lysine (XI), m 130°. Trypsin attacked III, V and X, the extent of cleavage indicating that only the α -linkage was broken. Trypsin kinase attacked V and XI, the latter only to 25%. Here again it appears that only the α -peptide linkage is broken. The 107% cleavage of VIII is considered questionable. All of these derivs were hydrolyzed more or less by N alkali at 37°. The important observation is that the ϵ -peptide linkage of lysine polypeptides is resistant to enzymes. A typical illustration is the cryptic cleavage of III and non-cleavage of VII. These observations support the view that the lysine component of proteins is linked only through its α -NH₂. A W DOX.

Further studies on the behavior of polypeptides containing proline toward the enzyme trypsin-kinase complex. EMIL ANDERHJELDEN and OTTO ZUMSTEIN. *Fermentforschung* 12, 341-9 (1931), cf *C A* 25, 77.—The proline peptides were prepd by coupling 1,4-dibromovaleryl chloride with the amino acid or dipeptide and closing the proline ring by amination. The above acid chloride was obtained by condensing $\text{CH}_3(\text{CO}_2\text{Et})_2$ with $(\text{CH}_3)_3\text{Br}$ and brominating, sapon. the $\text{Br}(\text{CH}_2)_3\text{CHBrCH}(\text{CO}_2\text{Et})_2$ with HBr , extg. with Et_2O and distg., and repeating the operation several times with the residue. dl-Prolyglycine (I), m 222-3°, dl-prolyl-d-alanine (II), m p not stated, dl-prolyl-d-leucine (III), m 225-6°, dl-prolylglutamic acid (IV), dl-prolyl-d-leucylglycine (V), m 245° (decompn) III + dl $\text{Me}_2\text{CHCH}_2\text{CHBrCOCl} \rightarrow$ viscous resin \rightarrow dl-leucyl-d-prolyl-d-leucine (VI), m 255-6° (decompn). Enzyme expts indicate that polypeptides which contain no NH, but rather an NH as in proline require a special enzyme complex for cleavage. Trypsin does not attack I or VI; it attacks II and IV only slightly, but III and V more readily. Trypsin kinase did not hydrolyze any of the proline peptides. III and VI are resistant to N alkali while I undergoes 50% cleavage in 2 hrs., then reaches an equib. II and IV are slowly hydrolyzed, V more readily. The expts show that enzymes do not merely accelerate hydrolyses already in progress but initiate reactions not otherwise occurring. The enzyme-substrate complex can behave differently from the substrate alone toward a given H ion concn. A W DOX.

The isomerization of E. Fischer's "carbethoxydiglycylglycine ester." M V CERCHIZ. *Bull soc chim* [4], 49, 324 8 (1931), cf *Ber* 36, 2074, 2106 (1903).—Fischer prepd carbethoxydiglycylglycine ester, $\text{EtO}_2\text{CCH}_2\text{NHCOCH}_2\text{NHCOCH}_2\text{NHCOCH}_2\text{NHCO}_2\text{Et}$ (I), m 160-1°, from $\text{EtO}_2\text{C}_2\text{NHCH}_2\text{CONHCH}_2\text{COCl}$ and $\text{NH}_2\text{CH}_2\text{CO}_2\text{Et}$ in CHCl_3 and by the action of $\text{ClCOCH}_2\text{NHCO}_2\text{Et}$ on $\text{NH}_2\text{CH}_2\text{CONHCH}_2\text{CO}_2\text{Et}$. On sapon. I formed first the acid ester, $\text{HO}_2\text{CCH}_2\text{NHCOCH}_2\text{NHCOCH}_2\text{NHCO}_2\text{Et}$ (II), m 208-10°, and then the diacid, diglycylglycinecarboxylic acid, $\text{HO}_2\text{CCH}_2\text{NHCOCH}_2\text{NHCOCH}_2\text{NHCO}_2\text{H}$ (III), m 206° (decompn). I found that the esterification of III gave an ester (IV) with the compn of I, but m 146-8°, and considering isomerization to have taken place, designated I and IV as α - and β -forms. Leuchs and La Forge (*C A* 2, 3078) held the opinion that they could be regarded as the tautomers, $\text{EtO}_2\text{CCH}_2\text{NHCOCH}_2\text{NHCOCH}_2\text{NHCOCH}_2\text{NHCO}_2\text{Et}$ and $\text{LiO}_2\text{CCH}_2\text{N}(\text{C}(\text{OH})\text{CH}_2\text{N}(\text{C}(\text{OH})\text{CH}_2\text{NHCO}_2\text{Et}))$. The possibility of carbamidodiacetic acid $\text{OC}(\text{NHCH}_2\text{CO}_2\text{H})_2$, m 205-10°, and its di Et ester, m 147-8°, being identical with III and IV was disposed of when the latter compds were prepd according to the original directions. Analysis showed that III was a diacid, $\text{C}_{11}\text{H}_{16}\text{O}_8\text{N}_4$, and mixed m p detns demonstrated the non identity of the pairs of compds in question. C R ADDINALL.

Hydroxy aldehydes. IX. Synthesis and properties of a 6-hydroxy aldehyde. BURCKHARDT HELFERICH AND GEORG SPARNBERG *Ber* 64B, 104-9(1931), cf *C A*, 19, 2931.—It was shown in earlier papers that 4 and 5-HO aldehydes exist, in large part, at least, as lactols (i. e., in a cyclic form) which on acetalization with acid MeOH give lactolides (cyclohemiacetals). Two derivs or forms of glucose are known in which a 1,6 lactole formation may be assumed, viz levoglucosan and Pringsheim's remarkably stable "γ glucose". It seemed of interest, therefore, to study a 6-HO aldehyde with respect to its ability to form a lactol and lactolides. The prepn of such an aldehyde from 1-methylcyclohexene by O_2 cleavage and subsequent partial reduction did not prove feasible as the ozonide is very explosive and could be prepd only in small quantity and in dil soln. O_2 cleavage of dihydrogeraniol (rhodinol), which is present along with its isomer, $HOCH_2CH_2CHMe(CH_2)_3C(CH_3)_2Me$, in natural citronellol and hydrogenated geraniol, gave no better result. O_2 cleavage of acetyl dihydrogeraniol (I) was somewhat more satisfactory, but here too, it was not possible to obtain with certainty the pure (+)-4-methylhexan-6-ol acetate (II), presumably because the *dl* compd was used as the starting material, for pure, sterically quite homogeneous (+)-citronellol (III) yielded the desired II. The free HO aldehyde (IV), obtained from II by catalytic sapon with NaOMe, is a very viscous liquid of faint odor which distills *in vacuo* without decompn, reduces hot Fehling soln and is sol to the extent of 1% in H_2O at room temp as detd by the Bertrand method. Its *n* agrees with the value calcd for the lactol form, but a very dil aq soln gives a color with fuchsin SO_2 as rapidly as an equimol soln of AcH in H_2O . Its rotation in $CHCl_3$ is much higher than that of the intermediate open form dihydrogeraniol, I and II. The free IV is apparently a desmotropic mixt of the open HO aldehyde and the lactol form which in the concd substance and in $CHCl_3$ lies far toward the lactol, in very dil H_2O soln toward the HO-aldehyde form. It gives in very good yield an easily volatile, strongly rotatory Me lactolide (V) of peppermint odor which does not reduce Fehling soln and is rapidly and completely hydrolyzed to IV by aq N HCl. Acetate of III (prepd from a III with n_D^{20} 1.45678, d_4^{20} 0.8568, $[\alpha]_D^{20}$ +1°, n_D^{20} 1.44383, d_4^{20} 0.8901, $[\alpha]_D^{20}$ 2.5° II (41% yield), b_{11} 112-3°, n_D^{19} 1.43675, d_4^{19} 0.9892, $[\alpha]_D^{19}$ 1.7°, soly. in H_2O 1.5%, gives no color with 1:1 NaOH even on long standing, but quite rapidly turns yellow in concd H_2SO_4 . IV (yield, 30%), b_{11} 103-6°, n_D^{19} 1.46383, d_4^{19} 1.0204, mol wt in AcOH 125.5-127, $[\alpha]_D^{19}$ 14.8° (no solvent), $[\alpha]_D^{19}$ 27.5-7.7° ($CHCl_3$), unchanged by NaOH, rapidly turns yellow and brown in H_2SO_4 , can be titrated quantitatively with Fehling soln. V (yield, about 40%), faintly yellow, b_{11} 35-41°, n_D^{17} 1.43549, d_4^{17} 0.9370, $[\alpha]_D^{17}$ 110.6°, mol wt in AcOH 141. C A R.

The constitution of so-called cyanoacetoacetic ester and a disputed synthesis of citric acid. GEORGES FAVREL AND CHARLES PREVOST *Bull. soc. chim* [4], 49, 243-61 (1931), cf *Compt. rend* 111, 647(1890).—The prepn of the Et monochloroacetyl acetate of Iffaler and Held (*Ann. chim* 23, 145(1891)), considered by them to be $ClCH_2COCH_2CO_2Et$, was repeated and the fraction b 188.9° and contg 21-38% of Cl was

isolated. On treatment with $(NH_4)_2CS$, only the thiazole, $MeC(CO_2Et)N(CNH_2)_2S$, m 174-5°, characteristic of the α Cl deriv could be sepd. On diazotization only the hydrazone derived from $AcCHClCO_2Et$ (I) was formed and F and P conclude that the ester of H and H is composed entirely of I. The CN deriv (II) of I, prepd according to the directions of H and H, was identical with the compd made by the action of aq KCN on the true α Cl deriv (*Ber* 23, 2339(1890)) whereas treatment of $NCCH_2COCH_2CO_2Et$ with KCN gave di Et succinylsuccinate,

$EtO_2CCH(COCH_2CH_2COCHClCO_2Et)_2$, m 127-8°. From the products of the reaction of alc HCl with II, 50% of I was recovered together with a non-volatile portion, free from $CO(CH_2CO_2Et)_2$, which on boiling with HCl for 2 hrs yielded $MeC(CO_2H)(OH)CHClCO_2H$, m 105-6°. This diacid was also obtained by the action of KCN and HCl on I, its formation being similar to that of $MeC(CO_2H)(OH)CH_2CO_2H$ from the reaction of KCN and HCl with $AcCH_2CO_2Et$ (*Ber* 24, 2344(1891)). The

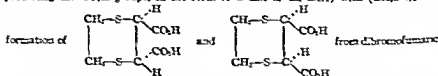
formula $MeC(CN)CH(CO_2Et)O$ (III) is assigned to II, and its formation is explained by the assumption of an intermediary cyanohydrin and consequent splitting out of HCl. By the action of the calcd amount of aq KCN on I, agitation of the soln with air for 1 hr per g mol of material, extrn with Et_2O and distn under reduced pressure, yields in excess of 50% of III, b_{10} 108°, were obtained. By treatment of 50 g of III

with 100 cc. of 4 N KOH for several min., addn. of a slight excess of 4 N H₂SO₄ and heating to 90°, MeC(CN) OH CH₂(OH)CO₂H, m. 155° (decomp.) was formed. By hydrolysis of III with KOH at 70° until 98% of the theoretical amt. of NH₃ was evolved, and neutralization with H₂SO₄, a viscous non-cryst. non-decolorable residue resulted which was divided. One portion was decomposed by heat into CO₂, H₂O, FeCl₂O and Me₂CO. The remainder, treated in MeOH with 10 vols. of Et₂O, gave an amorphous hygroscopic ppt. m. 110°, identified by its IR and Ag salts, its reactions with CaSO₄, FeCl₂, resorcinol and 3 CuH₂O, and by comparison with a synthetic product, as MeC(CN) OH CH₂(OH)CO₂H. It is concluded that III can lead only to methyltartronic acid deriva. and not to deriva. of acetonecarboxylic acid as maintained by H and H₂, that though it is conceivable that III might add HCN to give deriva. of MeC(CO₂H)CH₂(OH)CO₂H or MeC(CO₂H) OH CH₂(CO₂H), i. e., deriva. of citric acid, such isomers have not been isolated in this repetition of the original synthesis. C. R. ARONOFF.

Extension of Michael's reaction. P. C. GUPTA AND M. N. CHAKRABARTY. *Proc. 12th Indian Sci. Cong.* 1928, 150.—Michael's reaction takes place between Na deriva. of AcCH₂CO₂Et, malonic ester and NCCN CO₂Et and compounds of the general formula RCH CHX or RC CX in which R is an org. radical and X is a strongly neg. radical like CO₂H or CN. G. and C. studied the action of the same Na deriva. on mustard oils and isocyanates. The compds., e. g., RNNaC(S CH₂CO₂Et), formed from mustard oils are easily decomposed by acids to yield compds. of the general formula RNH₂CO₂Me together with CO₂ and alc., isocyanates form compds. of the type RNHCOMe. No details are given. E. J. C.

Condensation of ortho esters with acetoacetic ester and malonic ester. F. T. SALT. *J. Am. Chem. Soc.* 53, 1826-9, (1931).—MeC(OEt)₂ (1.2 g) and AcCH₂CO₂Et with 2 mols. AcO, boiled 4 hrs., gave 9 g. of *Et mesodichloromalonate*, yellowish white greek-thick liquid, b. 110-23° (decomp.), the low yield is probably due to partial decomp. during distn., on standing the color deepened, addn. of ZnCl₂ during the condensation period produced no favorable effect upon the yield, the Ca salt, CaH₂O₂Ca, greenish blue, is characteristic, CaH₂N gives a violet color; warm H₂O hydrolyzes the salt to CaO. CH₃(CO₂Et)₂ gives 3 g. of *Et mesodichloromalonate*, greek-thick liquid, b. 125-6°, Ca salt, greenish blue. EtCH₂CO₂Et (7, b. 125-30°, b. 240°, d₄²⁰ 1.0225, n_D²⁰ 1.408 (124 g) and AcCH₂CO₂Et gave 40 g. *Et picrodichloromalonate*, b. 135-6°, d₄²⁰ 1.0449, n_D²⁰ 1.5312, addn. of ZnCl₂ increased the yield slightly, Ca salt, light blue. I (75 g) and CH₃(CO₂Et)₂ gave 25 g. of *Et picrodichloromalonate*, b. 180-90°, d₄²⁰ 1.0738, n_D²⁰ 1.4920, Ca salt, light blue, ZnCl₂ increased the yield considerably. C. J. WEST.

α,α'-Diisomerism in closed-ring compounds having two sulfur atoms in the nucleus. CHITTARANJAN BARI. *Proc. 12th Indian Sci. Cong.* 1928, 155.—Dibromosuccinic acid occurs in two different isomeric forms according as it is obtained by brominating fumaric or malonic acid. The assumption, therefore, that they still retain their steric characteristics, is justifiable. If now they are made to condense (after protecting the CO₂H groups, in the form of a salt or an ester) with (CH₃SH)₂, the



and dibromomalonic acids, resp., may be expected. The condensations have actually been effected in the presence of NaOEt in alc. suspension (the mercaptan working best as an alkali mercaptide). The products of condensation, though different, have not yet been obtained in a sufficiently pure form. E. J. C.

The interaction of sulfonic acid and bromine on ethyl benzylideneacetacetate. B. M. GUPTA AND S. C. ROY. *Proc. 12th Indian Sci. Cong.* 1928, 158.—With a view to obtaining α-Br deriva. of variously substituted Et glutarates, Et benzylideneacetacetate was treated with concd. H₂SO₄ and Br₂. A cryst. compd. m. 159° having the empirical formula C₁₁H₁₁BrO₄ was obtained. By studying the products of hydrolysis and oxidation, the constitution of the compd. has been fixed as mono-Et α,α'-diaceto-α-bromo-β-phenylglutarate. This formula has been further confirmed by the study of the product of its interaction with aniline, which is a cryst. compd. having the empirical formula C₁₈H₁₅NO₄ and m. 131°. E. J. C.

cis trans-Isomerism in ethyl carbethoxythiocarbamate. Synthesis of four-, five-, six- and seven-membered heterocyclic compounds from ethyl carbethoxythiocarbamate. P. C. GUHA AND N. C. DUTT. *Proc 15th Indian Sci. Cong* 1928, 157; cf. *C. A.* 24, 1113—Et carbethoxythiocarbamate has been isolated in 2 isomeric forms, one m 44° and the other 141°, and the latter has been found to be convertible into the former by boiling with water contg. a drop of HCl. With $(CH_3NH_2)_2$, the lower-melting variety gives $NHCO N C(OEt) NH CH_2 CH_3$ (insol in alkali), whereas the higher-

melting variety gives $NH N C(SH) NH CH_2 CH_3$ (sol in alkali) Et carbethoxythiocarbamate has been made to react with aromatic amines, N_3H , $PhNHNH_2$, urea, phenylurea, thiourea, semicarbazides and α - $C_{12}H_{25}(NH_2)_2$, and thus a large no

of interesting heterocyclic compds have been obtained E. J. C

Oxidation of carbohydrates and fats by air in presence of yellow phosphorus. SACHINDRA NATH CHAKRABARTI AND N. R. DHAR. *J Phys Chem* 35, 1114-22 (1931), cf. *C. A.* 24, 349—P can be used as an inductor in the oxidation of carbohydrates and fats. Finely divided P was found to be most effective, colloidal P next and small pieces of P least. Colloidal P should have been most effective, but its action was retarded by the ale in the soln from which it was prepd. The order of oxidation of the carbohydrates and fats investigated is given by C. and D. The fact that starch is more easily oxidized than glucose shows that in slow photochem or induced oxidation, starch does not pass through the stage of glucose before its oxidation to CO_2 .

LOUISE KELLEY

Optical rotation and atomic dimensions. IX. Halogenotetraacetyl derivatives of mannose. Their configurational peculiarities. D. H. BRAUNS. *J Am Chem Soc* 53, 2004-5 (1931), cf. *C. A.* 23, 3111—Preliminary note, pointing out that mannose derives data from other monosaccharides studied in that ratios of the differences in sp rotation Cl , Br - Cl and I - Br (41 23 35) are not the same as the ratios of the corresponding differences in at diameters. It is suggested that for the mannose derives the halogens on the 1 hand and the atoms of the 2nd C atom on the other influence each other.

C. J. WEST

Unsaturated reduction products of the sugars. XV. Structure of the pseudoglucals. MAX BERGMANN AND WERNER FÄBUDENBERG. *Ber.* 64B, 158-61 (1931); cf. *C. A.* 24, 6027—The velocity of cleavage of glucosides is influenced by the HO groups of the sugar components. The chemistry of the pseudoglucals affords valuable information on this influence. Diacetyl ψ -glucal (I) on hydrogenation of the double bond gives an aldehyde which does not reduce Fehling soln and it was therefore assigned the structure $AcOCH_2CH(OAc)CH(OH)CH_2CH_2CH_2OH$. It is converted by

$HCl(OMe)_2$ into its α -Me lactolide (II) which with $Ba(OH)_2$ at room temp. gives ψ -glucal α -Me lactolide (III). III is also obtained with emulsin from the Me lactolide of ψ -cellobial (IV). It follows that in ψ -glucal and IV the unsatd sugar residue has the same structure and IV is 4 glucosido- ψ -glucal. Since in IV the glucosido residue is on C atom 4, the furoid structure for the unsatd. part of the mol is excluded. Furthermore, I, as well as III and the pentaacetate of IV, give $(CHIO)_2$ (identified as the phenyllosazone) with O_2 , showing that the double bond in these glucals lies between C atoms 2 and 3 and fully confirming the above structures. As the lactolides are converted by catalytic hydrogenation into the lactolides of 2,3 bisdesoxyglucose and 2,3-bisdesoxycellobiose, these latter must also have a pyroid structure. All these lactolides, highly sensitive to very dil. acids, have the same ring structure as α - and β -Me glucoside, and their sensitivity is in full agreement with the theory that it is due to the absence of HO groups on C atoms 2 and 3. II, b_p , 110°, n_D^{20} 1.4570, $[\alpha]_D^{25}$ 143.0° ($C_{11}H_{14}$), rapidly decolorizes Br. III (67% yield), b_p , 118-20°, n_D^{25} 1.4915 $[\alpha]_D^{25}$ 102.3° (H_2O), does not reduce Fehling soln but rapidly decolorizes Br.

C. A. R.

Synthesis of a tetrasaccharide acetate. BURCKHARDT HELFEBICH AND RICHARD GOOTZ. *Ber.* 64B, 109-14 (1931)—The improvement in the method of coupling acetohalogenosugars with tetraacetylglucose (*C. A.* 24, 3762) has made it possible to prep a synthetic tetrasaccharide in somewhat greater yield than had hitherto been obtained and to study it somewhat more thoroughly. Gentiosidogenotriose (I) was chosen in order to have a uniform chain of 4 glucose mols. β Gentiosidogenotriose d-glucose hendecaacetate (II) was converted, by specially cautious treatment with $AcOH$ - $IIBr$, into the acetobromo compd. (III) which, when coupled with 1,2,3,4- β tetraacetyl-d-glucose (IV),

gave the desired *5-gentiobiono-3-gentiobiose tetradecaacetate* (V), in well-crystd. form. Cleavage of the Ac groups from both the tri- and the tetrasaccharide has thus far given only amorphous sugars easily sol. in H_2O . The purity of the V was established not only by analysis and Ac detns. but also by titration according to Willstätter-Schudel. Its soly. in 100 cc. alc. at 0° is 0.049 g., but only when it is very pure, while a product which has been recrystd. 8-12 times gives satisfactory results on analysis and mol. wt., Ac and rotation measurements, the above low soly. is obtained only after 20 crstns. This shows the difficulty of sepg. and purifying such sugar acetates and points to the necessity for caution in judging of their homogeneity, especially when dealing with products of the partial hydrolysis of polysaccharides. II, obtained in 50% yield from acetobromogentiobiose and IV in $CHCl_3$ with $CaCl_2$, Ag_2CO_3 and I, m. $219-20^\circ$, in ice-cold $CHCl_3$ with $AcOH$ acid. at 0° with HBr it gives 61% III, m. $193-4^\circ$ (cor.), $[\alpha]_D^{25} -0.3-3.3$ ($CHCl_3$). V, sinters around 135° , resolidifies and m. $277-9^\circ$ (cor.), $[\alpha]_D^{25} -11.1^\circ$ ($CHCl_3$), mol. wt. (titration according to Willstätter-Schudel) 1200-36, mol. red.-ing. power (Bertrand) (glucose = 100) 97. C. A. R.

Reaction relating to carbohydrates and polysaccharides. XXIV. The constitution of levan and its relation to inulin. HAROLD HIRSHAY, R. STUART TIRSON AND F. BRADY. *Can. J. Research* 4, 221 (30'1931), cf. C. A. 25, 723.—A description is given of the properties of levan, a polysaccharide obtained by the action of *Bacillus mesentericus*, and of its enzyme, on sucrose, in particular, its complete hydrolysis to fructose, and behavior when subjected to acetylation and methylation. By hydrolysis of trimethyllevan, a cryst. trimethylfructose was obtained. This yielded a trimethylmethyl fructose on condensation with $MeOH$ contg. HCl . The properties of these compds. are described for the first time. On methylation and subsequent hydrolysis of the glucide, a tetramethylfructose was obtained having properties identical with those of the tetramethylfructose obtained by hydrolysis of heptamethylsucrose. Oxidation of the new trimethylfructose yielded a dibasic dimethylactol acid, and esterification of the latter, followed by methylation and treatment with dry NH_3 , gave a cryst. diamide. These results show that the new cryst. trimethylfructose is 1,3,4-trimethylfructofuranose. It follows from this that levan is a polymerized anhydrofructofuranose with linkages at positions 2 and 6 of the fructose chain, thus differing from inulin which has been shown to be an anhydrofructofuranose with linkages at positions 1 and 2. The synthesis of levan thus represents a further example of the polymerization of an anhydro sugar taking place through the medium of the $\gamma, \delta-HO$ groups. J. W. S.

Carbon rings. XVI. Relationship between density and molecular arrangement in a homologous series of normal aliphatic and cyclic hydrocarbons. M. STOLL AND G. STOLL-COENRE. *Helv. Chim. Acta* 13, 1185-1200 (1930), cf. C. A. 25, 922.—The article is a math. study of the characteristic d. curves of the normal cyclic hydrocarbons and ketones at 20° . Based on this study, S. and S.-C. build up a theory concerning the mol. structure of these compds. A. WIESEN

Competitive hydrogenations. II. F. F. DWORKY AND HOMER ADKINS. *J. Am. Chem. Soc.* 53, 1878-79 (1931), cf. C. A. 24, 94.—The rates of hydrogenation (listed in order of decreasing rate) of quinoline, C_6H_6 , $PhMe$, $PhOH$, $PhCH_2OH$, C_6H_5N , $PhNH_2$, $BzNH_2$, and $PhNH_2$, have been detd. over a Ni catalyst at $125-75^\circ$ and under pressure of H_2 from 125-200 atm. Thirty two mixts. of these compds., each mixt. contg. 2 components, were subjected to the action of H_2 over a Ni catalyst and the mixt. allowed to take up as much H_2 as would have been sufficient to completely hydrogenate 1 of the components of the mixt. The rate of absorption of the H_2 and its distribution between the components of the mixt. were detd. by analysis of the reaction mixt. It was found that if the compds. were listed in the series quinoline, C_6H_5N , $PhNH_2$, $PhCH_2OH$, $PhOH$, $PhNH_2$, C_6H_6 , $BzNH_2$, and $PhNH_2$, in any mixt. contg. 2 of them, the 1 higher in the series would receive the greater proportion of the H_2 . C_6H_5N and quinoline in mixts. with other compds. took all, or practically all, of the H_2 . Apparently there is no relationship between the relative rates of hydrogenation of these compds. singly and in mixts. The relative rates of hydrogenation of the components of a mixt. are at least in part presumably dependent upon selective absorption by the catalyst through 2 possible acceptors of H_2 . The hydrogenation of $PhNH_2$ was accelerated by $PhMe$ and C_6H_6 and the hydrogenation of $PhNH_2$ by the presence of $PhOH$ in the reaction mixt. The ratio of cyclohexylcarbamol to $PhMe$ produced in the hydrogenation of $PhCH_2OH$ was increased by the presence of $PhOH$ or $PhNH_2$ and especially of C_6H_6 or $BzNH_2$ in the reaction mixt. The original should be consulted for the table of data. C. J. WEST

Syntheses in the cyclobutanol series. L. BLANCHARD. *Bull. soc. chim.* [4], 49,

279-309(1931), cf *C A* 21, 3888—The ethers of the dihalohydrins, $\text{XCH}_2\text{CH}(\text{OR})\text{CH}_2\text{X}$, on condensation with $\text{NaHC}(\text{CO}_2\text{Et})_2$ gave $\text{XCH}_2\text{CH}(\text{OR})\text{CH}_2\text{CH}(\text{CO}_2\text{Et})_2$, which was saponid to $\text{XCH}_2\text{CH}(\text{OR})\text{CH}_2\text{CH}(\text{CO}_2\text{H})_2$. Traube and Lehmann (*Ber* 34, 1977(1901)) have shown that the acidification of the condensation product of

$\text{ClCH}_2\text{CHCH}_2\text{O}$ with $\text{NaHC}(\text{CO}_2\text{Et})_2$ gives a lactone, $\text{ClCH}_2\text{CHCH}_2\text{CH}(\text{CO}_2\text{Et})\text{COO}$

Compds of the type $\text{CH}_2\text{CH}_2\text{CH}_2\text{C}(\text{CO}_2\text{Et})_2$ formed by malonic ester condensation with $\text{Cl}(\text{CH}_2)_2\text{Br}$ or $\text{Br}(\text{CH}_2)_2\text{Br}$ are well known. By the use of $\text{XCH}_2\text{CH}(\text{OR})\text{CH}_2\text{X}$,

it would be possible to evade lactonization and so to prep $\text{ROCHCH}_2\text{CH}_2\text{CH}_2\text{C}(\text{CO}_2\text{Et})_2$,

from which $\text{HOCHCH}_2\text{CH}_2\text{CH}_2\text{C}(\text{CO}_2\text{H})_2$ (I) might be prepd. The removal of the R group of the ether might destroy the cyclobutane ring, though this difficulty would

be avoided by synthesis of the corresponding formal, $\text{ROCH}_2\text{OCHCH}_2\text{CH}_2\text{C}(\text{CO}_2\text{Et})_2$, which could be hydrolyzed by hot H_2O to I. Furthermore, the prepn. of the formals $\text{ROCH}_2\text{OCH}(\text{CH}_2\text{X})_2$ is more convenient than that of the dihalohydrin ethers. For-

mals 1 *Mixed formals*—The addn of XCH_2OR to $\text{ROCHCH}_2\text{CH}_2\text{O}$ always yields mixed formals of the type $\text{ROCH}_2\text{CH}(\text{OCH}_2\text{OR})\text{CH}_2\text{X}$. By the addn of MeOCH_2Cl , EtOCH_2Cl and AmOCH_2Cl to epichlorohydrin the following mixed formals were formed. *Me 1,3-dichloroisopropyl*, $\text{MeOCH}_2\text{OCH}(\text{CH}_2\text{Cl})_2$, b_{11} 80-1°, d_{11} 1.237, n 1.45412, M R. 37.88, calcd. 38.31, *Et 1,3-dichloroisopropyl*, b_{11} 90-1°, d_{11} 1.182, n 1.44912, M R. 42.44, calcd. 42.93, *Am 1,3-dichloroisopropyl*, $\text{AmOCH}_2\text{OCH}(\text{CH}_2\text{Cl})_2$, b_{11} 133-5°, d_{11} 1.09, n 1.4506, M R. 56.55, calcd. 56.36. On addn. of EtOCH_2Br , AmOCH_2Br and EtOCH_2I , to epichlorohydrin, chlorobromo and chloriodo formals were produced, *Et 1,3-chlorobromoisopropyl*, b_{11} 110-2°, d_{11} 1.409, n 1.46954, M R. 45.81, calcd. 45.82, *Am 1,3-chlorobromoisopropyl*, b_{11} 142-4°, d_{11} 1.277, n 1.48856, M R. 59.52, calcd. 59.63, *Et 1,3-chloriodoisopropyl*, $\text{EtOCH}_2\text{OCH}(\text{CH}_2\text{Cl})\text{CH}_2\text{I}$, b_{11} 124-5°, d_{11} 1.6528,

n 1.50882, M R. 50.28, calcd. 50.80. MeOCH_2Cl added to $\text{AmOCH}_2\text{CHCH}_2\text{O}$ yielded *Me 1,3-chloroamylisopropyl formal*, $\text{MeOCH}_2\text{OCH}(\text{CH}_2\text{Cl})\text{CH}_2\text{OAm}$, b_{11} 118°, d_{11} 1.01, n 1.43587, M R. 58.10, calcd. 58.17. 2 *Symmetrical formals*—A mixt. of dichlorohydrin and its ClCH_2 ether reacts slowly at room temp, evolving HCl with formation of a sym. formal. The elimination of HCl is favored by the presence of $\text{Mg}(\text{OH})_2$. *Dichlorohydrin formal*, $\text{CH}_2[\text{OCH}(\text{CH}_2\text{Cl})_2]_n$, m . 51°, *chlorobromohydrin formal*, m . 54-5°; *chloriodohydrin formal* m . 60°, *dibromohydrin formal*, m . 68-9°. The condensation of $\text{ClCH}_2\text{CH}(\text{OMe})\text{CH}_2\text{Br}$ with $\text{NaHC}(\text{CO}_2\text{Et})_2$ formed *di-Et 1-chloro-2-methoxybutane-4,4-dicarboxylate*, b_{11} 157°, d_{11} 1.135, n 1.44682, M R. 63.11, calcd. 62.75. The sapon of this ester gave an acid that was very difficult to isolate in a pure state. The condensation of $\text{BrCH}_2\text{CH}(\text{OAm})\text{CH}_2\text{Br}$ with $\text{Na}_2\text{C}(\text{CO}_2\text{Et})_2$ was carried out in 2 stages and gave *di-Et amyloxycyclobutanedicarboxylate* (II), b_{11} 175°, d_{11} 1.011, n 1.4436, M R. 75, calcd. 74.22. By digesting with 30% KOH for 5 hrs., the free acid (III) was prepd., which in turn was converted into its Cu salt, *diacid chloride*, b_{11} 143-5°, *diamide*, m 171.5°, *amyloxycyclobutanecarboxylic acid*, m 222-3°, *dianilide*, m 175°. By heating to 120°, III gave off CO_2 and formed *amyloxy-*

cyclobutanecarboxylic acid, $\text{AmOCHCH}_2\text{CH}_2\text{CHCO}_2\text{H}$, b_{11} 164-6°, d 1.003, n 1.45412,

M R. 50.21, calcd. 49.36; *amide*, m 131-2°. $\text{AmOCH}_2\text{OCHCH}_2\text{CH}_2\text{C}(\text{CO}_2\text{Et})_2$, prepd. similarly to II, was saponid to the diacid by the use of gaseous HCl at low temps. to prevent the hydrolysis of the formal group. By boiling the diacid with water for

5 hrs., cleavage into the *diacid alc.*, $\text{HOCHCH}_2\text{CH}_2\text{C}(\text{CO}_2\text{H})_2$, m 125°, and the formal, $\text{CH}_2(\text{OAm})_2$, d_{11} 0.843, took place. I and a similar MeO compd. showed an exaltation of the M . R. and the paper is concluded by a discussion of the M . R. of cyclobutane compds.

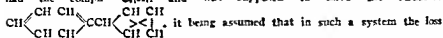
C. R. ADDINALL

Cyclohexylmagnesium chloride and bromide. HENRY GILMAN and E. A. ZOELLNER. *J. Am. Chem. Soc.* 53, 1945-8(1931)— $\text{C}_6\text{H}_{11}\text{MgCl}$ and $\text{C}_6\text{H}_{11}\text{MgBr}$ have been prepd. in 80-96% yields. In its reaction with $\text{Alg C}_6\text{H}_{11}\text{Br}$ differs significantly from other sec. bromides. General recommendations are given for the use of RCl in preference to RBr and RI compds. for the prepn. of Grignard reagents.

C. J. WEST

Dehydrogenation catalysis of bicyclic hydrocarbons. N. D. ZELINSKI and I. N.

TITZ *Ber* 64B, 183-8(1931)—Some time ago Z obtained by dehydrogenation of cyclohexylcyclopentane (I) a cryst. compd (II), m 104° , which apparently had the compn $C_{11}H_{12}$ and was supposed to have the structure



of H atoms from the 6-membered ring stimulated a similar loss from the 5-membered ring although the dehydrogenation of a cyclopentamethylene ring was so unexpected that the synthesis of compds. of the type I was undertaken to check these results (C A 21, 900). The supposed I had been synthesized from $\text{C}_{11}\text{H}_{12}\text{MgBr}$ (III) and cyclopentanone (IV) and as the reaction seemed to proceed normally there seemed to be no doubt that the product was I. More thorough study showed, however, that II is really Ph₂ and that the action of Mg on $\text{C}_{11}\text{H}_{12}\text{Br}$ in Et_2O yields chiefly $(\text{C}_{11}\text{H}_{11})_2$, the normal III which with IV gives the expected tertiary alc. being formed in only small quantity. Heating the reaction product with aq. $(\text{CO}_2\text{H})_2$ yielded, together with the $(\text{C}_{11}\text{H}_{11})_2$, only a small quantity of an unsatd hydrocarbon. Likewise, $\text{C}_{11}\text{H}_{12}\text{MgBr}$ and cyclohexanone yielded no appreciable amt. of the expected tertiary alc., and III with 3 methylcyclopentanone gave chiefly $(\text{C}_{11}\text{H}_{11})_2$. The true I was accordingly prepd. by hydrogenation of phenylcyclopentane (V), b 217° , d_4^{20} 0.9474, n_D^{20} 1.5280, which was obtained in 54% yield from $\text{C}_{11}\text{H}_{12}\text{Cl}$, $\text{C}_{11}\text{H}_{11}$ and AlCl_3 . Passed at the rate of 4 drops per min. in H_2 over Pt-charcoal at $220-5^{\circ}$, V gave I, b 214° , d_4^{21} 0.8780, n_D^{21} 1.4728, which, when passed 3 times at the rate of 2 drops per min. over Pt-charcoal at 300° , yielded V, b $215-7^{\circ}$, n_D^{21} 1.5192, with no higher boiling product and not a trace of crystals of Ph₂. This passivity of the pentamethylene nucleus in bicyclic systems toward dehydrogenation catalysis is further confirmed by the following expts. Phenylcyclopentylmethane (VI), b $233-5^{\circ}$, d_4^{18} 0.9345, n_D^{18} 1.5206 (obtained together with $(\text{PhCH}_2)_2$ and a little $(\text{C}_{11}\text{H}_{11})_2$ from PhCH_2Cl , $\text{C}_{11}\text{H}_{11}\text{Cl}$ and Na), remained unchanged when passed over Pt-charcoal at 300° . At 200° in H_2 it yielded cyclohexylcyclopentylmethane, b $225-7^{\circ}$, d_4^{18} 0.8789, n_D^{18} 1.4775, which, when passed over Pt-charcoal at 300° , regenerated the VI. C. A. R.

Influence of substitution on the oxidation of side chains in the benzene nucleus. P. S. VARMA AND P. B. PANICKER, *Proc 13th Indian Sci Cong* 1923, 150-1—Using a neutral soln. of KMnO_4 , V. and P. oxidized o-, m- and p- $\text{O}_2\text{NC}_6\text{H}_4\text{Me}$, o-, m- and p- $\text{MeC}_6\text{H}_4\text{CO}_2\text{H}$, o- and p- $\text{ClC}_6\text{H}_4\text{Me}$, o- and p- $\text{BrC}_6\text{H}_4\text{Me}$, o- and p- $\text{IC}_6\text{H}_4\text{Me}$, o-, m- and p- $\text{AcHNCC}_6\text{H}_4\text{Me}$ and m- and p- $\text{MeC}_6\text{H}_4\text{OMe}$ under identical exptl. conditions. Conclusion. In the p-series the order (the degree of oxidation decreasing) in which the side chain is oxidized is $-\text{CO}_2\text{H}$, $-\text{NO}_2$, $-\text{Br}$, $-\text{Cl}$, $-\text{OMe}$, $-\text{I}$, $-\text{NHAc}$ and in the o-series the order is $-\text{NO}_2$, $-\text{CO}_2\text{H}$, $-\text{Br}$, Cl , $-\text{NHAc}$ and $-\text{I}$. With regard to position isomers, there does not seem to be any regularity. p- $\text{NO}_2\text{C}_6\text{H}_4\text{Me}$ is oxidized most, the m-compd. the least, whereas in the acetyltoluidines, the o-deriv. comes first and the p-deriv. last. No details are given. E. J. C.

Obtaining an intermediate form in an acetylenic transposition. M. BOURGUEL, *Compt. rend* 192, 686-8(1931), cf. C. A. 22, 2748.—Treatment of PhC, CMe (I) with NaNH_2 at $60-70^{\circ}$ for 3 hrs. followed by cold very dil. acid yielded true phenylpropine, $\text{PhCH}_2\text{C} \equiv \text{CH}$ (II), and a new hydrocarbon, b₁₁ 72° , d_{11} 0.9385, n_D^{20} 1.582, believed to be $\text{PhCH} \equiv \text{CCH}_2$ (III). This does not react with CuCl , is not hydrated by H_2SO_4 , and its Raman effect is different from I or II. Further treatment of III with NaNH_2 yields II. II treated with dry KOH at room temp. yields first III, then I. III polymerizes at 60° quickly to a soft resin. It forms a diketone, PhCOCOMe , either by taking up O_2 from the air, or by the action of $\text{Hg}(\text{OAc})_2$. L. E. GILSON.

Physical properties of various nitro derivatives. LOUIS DESVERGNES, *Rev chim ind* 40, 34-7(1931), cf. C. A. 21, 3607, 23, 4207, 24, 350.—The solidification point of 2,4,6-(O_2N)₃ $\text{C}_6\text{H}_3\text{Me}$ on 3 detns was 80.83° . Its soly at 19° in 100 g of AcOEt is 30.921, abs. LiOH 1.420, MeOH 2.944. The soly of tetral at 18° for 100 g of solvent is AcOEt 13.916, Me_2CO 64.111, abs. EtOH 0.537, MeOH 1.320, C_6H_6 5.070, $\text{C}_6\text{H}_5\text{N}$ at 19.5° 43.997, C_6H_6 at 19.5° 3.137. Tetral is decomposed by $\text{C}_6\text{H}_5\text{N}$, 2,4-(O_2N) $\text{C}_6\text{H}_3\text{NHMe}$, recrystd. from Me_2CO , m (3 detns.) 176° , $175.5-6^{\circ}$, 176° , its soly in H_2O at 18° is 0.002%, at 59° 0.004%, at 100° 0.056%. Its soly at 17° in 100 g of solvent is AcOEt 0.718, Me_2CO 1.064, 96% EtOH 0.079, abs. EtOH 0.116, MeOH 0.111, C_6H_6 0.319, CHCl_3 at 17.5° 0.699, anhyd. EtOH at 17.5° 0.108, $\text{C}_6\text{H}_5\text{N}$ at 17.5° 3.454, CS_2 at 18.5° 0.051, CCl_4 at 18.5° 0.024, C_6H_6 at 18.5° 0.276, 2,4,6-(O_2N) $\text{C}_6\text{H}_3\text{NH}_2$ m (3 detns.) $190.5-1^{\circ}$, $190.5-1^{\circ}$, $190.5-1.5^{\circ}$. Its soly in H_2O at 19.5° is

0.106, at 50° 0.012, at 100° 0.123. Its soly in 100 g of solvent at 19.5° is AcOEt 1.89%, Me₂CO 4.70%, 96% EtOH 0.120, abs EtOH 0.127, MeOH 0.245, C₂H₅ 0.907, CHCl₃ at 17° 0.322, anhyd Et₂O at 17° 0.121, C₂H₅N at 17° 13.638, CS₂ at 17.5° 0.013, CCl₄ at 17.5° 0.003, C₂H₄ at 21.5° 1.108, 1,3,4,5-C₄H₃(NO₂)₄, m (3 detns) 126.1°, 126.2°, 126.2°. This product is decomposed by H₂O with formation of picric acid. Its soly in 100 g of solvent at 23° is C₂H₅ 2.291, CHCl₃ 0.577, CS₂ 0.094, CCl₄ at 25° 0.010. This product is decomposed by C₂H₅N and C₂H₄. Tetranitropentacerythritol m (3 detns) 140°, 139.5°, 139.5°. Its soly in H₂O at 19° is 0.010, at 50° 0.010, at 100° 0.035. Its soly in 100 g of solvent is AcOEt at 19° 6.322, at 50° 17.860, Me₂CO at 19° 25.632, at 50° 57.960, 96% EtOH at 10° 0.032, abs EtOH at 21.5° 0.120, MeOH at 21.5° 0.324, C₂H₅ at 19° 0.215, CHCl₃ at 19° 0.050, anhyd Et₂O at 19° 0.105, C₂H₅N at 19° 5.436, at 50° 8.567, CS₂ at 17.5° 0.014, CCl₄ at 17.5° insol. C₂H₄ at 17.5° 0.623.

P. THOMASSET

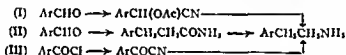
Nitration of benzene. P. S. VARMA AND K. A. JOSHI *Proc 15th Indian Sci Cong* 1928, 151—C₆H₆ was nitrated by means of HNO₃ alone in the presence of Zn, Cu, Sn and Sb as catalysts. It was possible to get some increase in the yield of PhNO₂ in the presence of metallic Sn but the yield so obtained can in no way be compared with the yield obtained by using a mixt of HNO₃ and H₂SO₄ as is done in the ordinary nitration. No details are given.

E. J. C.

Nitration of *m*-dichlorobenzene. H. S. JOIS AND B. L. MANJUNATH *Half-Yearly J. Mysore Univ* 4, 239-40 (1930)—*m*-C₆H₄Cl₂ (100 g) was nitrated according to the method of Nietzki and Schedler (*Ber* 30, 1666 (1897)) using HNO₃ of d 1.52. The mixt was kept at 80° for 30 min, then cooled and added to H₂O + ice. The ppt, crystd from EtOH, gave pure 4,6-dinitro-1,3-dichlorobenzene (I), rectangular plates, extinction about 40°, m 102.3°. The mother liquor was evapd to dryness, and the residue dissolved in Et₂O, washed with H₂O and dried, (Cl 30.3%, calcd for C₆H₂(NO₂)₂Cl₂, 29.9%). Dissolved in EtOH at 50°, it gave on cooling a crop of crystals of I. On boiling down and then cooling, this last mother liquor sometimes gave crystals of 2,4-dinitro-1,3-dichlorobenzene (II) (acicular crystals, extinction angle 10-12°, m. 70-1°), but it was not possible to det the exact exptl conditions for obtaining II. Recrystn from concd H₂SO₄ (cf. Dann, *C. A.* 24, 1376) gave II, m. 68-9°. The total yield was 126 g of I and 30 g of a mixture of I and II.

GEORGE CALINGAERT

New and improved methods in the formation of pharmacologically important amines. II. Synthesis of β -arylethylamines from aromatic aldehydes and carboxylic acids. KARL KINDLER *Arch Pharm* 269, 70-8 (1931), cf *C. A.* 21, 260-8-9—The following scheme indicates the methods followed in the present study in the prepn of β -arylethylamines:



Under I the synthesis of anisaldehyde, methylvanillin, piperonal and myristicinaldehyde was carried out, under II, the amides of *p*-methoxy-, 3,4-dimethoxy- and 3,4-methylenedioxyhydrocinnamic acid, under III, anisoyl chloride. The lowest yields obtained under I and II amounted to 70% of the theoretical. The amines obtained are convertible into tyramine, hordenine, ephrine, laudanoline, papaverine, hydrastrine, berberine and cotarnine. With W. PASCHKE Reduction expts with II under pressure in a specially constructed app are described in the prepn of homopiperonylamine from 3,4-methylenedioxyacetylmandelic nitrile, homoveratrylamine from 3,4-dimethoxyacetylmandelic nitrile, tyramine Me ether from *p*-methoxyacetylmandelic nitrile, C₇H₁₁O₂N, m 38-9°, b₁₀ 105°, and homomyristiclylamine from 3,4-methylenedioxy-5-methoxyacetylmandelic nitrile, tyramine Me ether from *p*-methoxybenzoylcyanide. With D. SCHMIDT Tyramine Me ether from *p*-methoxyhydrocinnamic amide, homoveratrylamine from 3,4-dimethoxyhydrocinnamic amide; homopiperonylamine from 3,4-methylenedioxyhydrocinnamic amide.

W. O. E.

Preparation of *p*-chloroaniline. K. HABIB HASAN *Proc 15th Indian Sci Cong* 1928, 158—Chattaway and Orton (1901) described a method of chlorinating AcNHPh with bleaching powder in glacial AcOH. A good yield of *o*- and *p*-comps is obtained. On hydrolysis of the anilides, however, the yields claimed for both anilines are not obtained, nor is the method of sepg *p*-chloroaniline suitable.

E. J. C.

Diazotization of difficultly diazotizable amines. S. KRISHNA AND RATAN LAL BHATIA *Proc 15th Indian Sci. Cong* 1928, 152—The usual methods of diazotization do not work with certain substituted amines, nor are such treatments as pressure, high

temp, concn, solvents, etc., applicable in all cases. Hence a method was worked out in which use was made of a basic solvent like pyridine. Make a soln of the desired amine in the min quantity of pyridine and add to it a strong soln of NaNO_2 (calcd quantity) in water. Add this mixt to a dil soln of HCl very slowly in small quantities. Diazotization proceeds satisfactorily. In this way tribromoaniline, 2,6-dichloro-*p*-nitroaniline, dibromo-*p*-aminobenzoic acid and dibromo-*p*-aminobenzaldehyde were successfully diazotized. **E. J. C.**

Substituted quaternary azonium iodides. V. The molecular state of phenyldimethylazonium, phenylmethylethylazonium, phenyldiethylazonium, phenylmethylbenzylazonium, phenylpropylbenzylazonium iodides in solution. HAWA KARTAR SINGH AND MOLA RAM SINGH. *Proc 15th Indian Sci Cong* 1928, 148, cf *C. A.* 15, 1506.—The mol. wts of these 5 compds are detd by using the Lumden Walker h. p. method, in water and EtOH . The solns employed are very dil, the concn, in most cases, ranging from 0.6 to 1.5%. **Conclusions** (1) The mol wts of all but phenylethylazonium iodide (I) are lower in EtOH than in water, in the case of I this relation is reversed. The degree of disson is, therefore, greater in EtOH , a solvent of low dielec. const., than in water, a solvent which has a very high value. It is, therefore, clear that the Nernst Thomson rule is not followed in these cases. (2) In H_2O , the degree of disson of these compds increases with increasing formula wt. of the base, but this regularity does not appear in EtOH . (3) The degree of disson of the salts increases with diln in both solvents. (4) The mol. wts furnished by the 2 methods, viz, the vol and the wt. methods, are nearly identical for very dil solns. No details or values are given. **E. J. C.**

Aryl chlorosulfonates and aryl sulfates. M. BATTAGAY AND L. DENIVILLE. *Compt rend* 192, 492-3 (1931).—Besides confirming the method for the synthesis of $(\text{PhO})_2\text{SO}$, b₁ 172-6°, by the action of SOCl_2 on NaOH in toluene at 0° (cf Voss and Blank, *C. A.* 25, 1797), B and D obtained a compd PhOSOCl , *Ph chlorosulfonate*, b₁ 94-6°. This is a new type of compd. It fumes in air, is decomposed by H_2O to PhOH , SO_2 and HCl , and by Zn to SO_2 and PhCl , and is the intermediate product in the formation of $(\text{PhO})_2\text{SO}$. **JANET E. AUSTIN**

Identification of phenols as the esters of 3,5-dinitrobenzoic acid. MAX PHILLIPS AND GEORGE L. KEENEY. *J. Am. Chem. Soc.* 53, 1924-8 (1931).—The 3,5-dinitrobenzoyl esters of phenols were prepd by refluxing 0.01 mol of the phenol and 0.01 mol of 3,5-(O_2N)₂ $\text{C}_6\text{H}_3\text{COCl}$ in 20 cc $\text{C}_6\text{H}_5\text{N}$ for 1 hr and pouring the cooled reaction product into 400 cc. cold 5% H_2SO_4 , then washing with H_2O and 2% NaOH . The m. ps are cor. The following data include the m. p. and values for α , β and γ , other crystallographic data are given in the original. 3,5-Dinitrobenzoyl deriv of PhOH , rods, 145.8°, 1.505, 1.690, >1.740, α - $\text{MeC}_6\text{H}_4\text{OH}$, plates, 138.4°, 1.490, —, >1.720, m - $\text{MeC}_6\text{H}_4\text{OH}$, plates, 165.4°, 1.475, 1.700, 1.740, p - $\text{MeC}_6\text{H}_4\text{OH}$, plates, 188.6°, 1.500, —, 1.740, 3,4- $\text{Me}_2\text{C}_6\text{H}_3\text{OH}$, rods, needles, 181.6°, 1.508, 1.670, >1.740, 2,6- $\text{Me}_2\text{C}_6\text{H}_3\text{OH}$, rods, plates, 158.8°, 1.512, 1.670, >1.740, 2,4- $\text{Me}_2\text{C}_6\text{H}_3\text{OH}$, rods, plates, 164.6°, 1.517, 1.645, >1.690, 3,5- $\text{Me}_2\text{C}_6\text{H}_3\text{OH}$, rods, 195.4°, 1.485, 1.650, >1.740, 3,6- $\text{Me}_2\text{C}_6\text{H}_3\text{OH}$, needles, 137.2°, 1.487, 1.655, 1.740, thymol, needles, 103.2°, 1.480, 1.625, 1.705, p -cyclohexylphenol, needles, 168.0°, 1.515, 1.625, 1.702, guaiacol, yellow needles, 141.2°, 1.530, —, >1.740, cresol, rods, needles, 170.6°, 1.550, 1.570, 1.640, 4,2- $\text{Pr}(\text{MeO})\text{C}_6\text{H}_3\text{OH}$, needles, 116.2°, 1.520, —, 1.650, eugenol, 130.8°, isoeugenol, 158.4°, α - $\text{C}_7\text{H}_7\text{OH}$, yellow needles, 217.4°, 1.490, 1.615, >1.740, β - $\text{C}_7\text{H}_7\text{OH}$, needles, 210.2°, 1.475, 1.740, >1.740. Details are given of the prepn. of cresol and 4,2- $\text{Pr}(\text{MeO})\text{C}_6\text{H}_3\text{OH}$. **C. J. WEST**

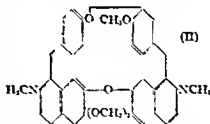
The pyrolysis of phenols. III. Influence of the time of heating and of gases present. AUGUST HAGEMANN. *Z. anorg. Chem.* 44, 231-4 (1931).— PhOH and pyrocatechol are passed through a hot tube in a current of N_2 , H_2 , CO_2 or steam. The % decompn. decreases as one passes from N_2 to steam, the yield of butadiene increasing in the same order from 33 to 56%. **GEORGE CALINGAERT**

Crotyl ethers of phenol. CHARLES D. HURD AND FRANK L. COHEN. *J. Am. Chem. Soc.* 53, 1917-22 (1931).—4-Chloro-2-hexene (10 g.), 7.0 g. PhOH , 40 cc. Me_2CO and 10 g. K_2CO_3 , refluxed 11 hrs, give 39% of α -ethylcrotyl *Ph* ether (I), b₁ 116-8°, and 1.8 g. α , α , β -dimethylcrotylphenol (II), b₁ 140-3°. Similarly, 4-chloro-2-heptene gives 65% of a propylcrotyl *Ph* ether (III), b₁ 153-4°, turns a dark red on cooling. 2,4-Dimethyl-2-chloro-3-pentene gives α , γ -trimethylcrotyl *Ph* ether (IV), very viscous, does not distil at 150° and 15 mm., n_D^{20} 1.2047. Heating I at 200-5° for 30 min gives 75% of II. III, heated at 228-30° for 30 min, gives 54% of α -methyl- β -ethylcrotylphenol, b₁ 154-5°. Heating IV at 160-70° for 30 min gives a mixt of PhOH and

2,4 dimethyl 1,3 pentadiene (33% yield) It is pointed out that the substituted allyl aryl ethers with the grouping $\text{ArOCH}_2\text{C} \equiv \text{C}$ or $\text{ArC}\equiv\text{CHRC} \equiv \text{C}$ may be expected to rearrange on heating whereas with a grouping such as $\text{ArOCR} \equiv \text{C}$ C pyrolysis leads to ArOH and a hydrocarbon C J West

Rearrangement of saturated alkyl phenyl ethers. Synthesis of isopropyl-phenol and -cresols. JOSEPH B NIEDERL AND SAMUEL NATIELSON *J Am Chem Soc* 53, 1928-34(1931) cf C A 25, 931 —For the conclusive proof of the reaction mechanism offered for the condensation of olefins with phenols under the influence of cored H_2SO_4 in the cold presented in the earlier paper, it was necessary to show that said alkyl Ph ethers would rearrange under the conditions used by Keenig (*Ber* 24, 3880(1901)) for this purpose the various iso-Pr Ph and t-olyl ethers were studied The ethers were prepd by heating 1 mol wt of the phenol, excess KOH (50 g) and excess (125 g) iso-PrBr (details given) the ethers were rearranged by adding 50 cc of the ether to 150 cc cored H_2SO_4 made up to 1 l with AcOH and refluxing 5 hrs PhOCH(CH₃)Me, b. 178° d₄²⁰ 0.973, n_D²⁰ 1.4992 rearrangement gives o-Me-CH(CH₃)OH, b. 213-4°, d₄²⁰ 1.012, n_D²⁰ 1.5315, iso-Pr o-t-ol ether b. 192° d₄²⁰ 0.953, n_D²⁰ 1.5040 2,4-Me-(Me-CH)₂C₆H₃OH, b. 252-4° d₄²⁰ 0.980, n_D²⁰ 1.5175 2,4,6-tri Br deriv. m. 223° iso-Pr m-tol ether b. 195° d₄²⁰ 0.931, n_D²⁰ 1.4959 3'-Me-Me-CH(CH₃)C₆H₃OH b. 230-3° d₄²⁰ 0.904, n_D²⁰ 1.5280 iso-Pr p-tol ether b. 84° b. 195° d₄²⁰ 0.927, n_D²⁰ 1.4932 4-methyl 2-isopropylphenol b. 234-7° d₄²⁰ 0.909, n_D²⁰ 1.5277 o-Me-CH(CH₃)C₆H₃OH-CH₂OH m. 130° 2-Methyl 4-isopropylphenol acid, from the K salt of the phenol and BrCH₂CO₂H, m. 87° 3-methyl 1-isopropyl deriv. m. 141° a 2nd deriv. m. 125-6°, was prepd from a condensation product, b. 95-100° 4-methyl 2-isopropyl deriv. m. 126° Proofs for the probable structure of these products are discussed and a general rule as to the way of migration of the alkyl groups is given As the possibility of rearranging even short chain, said Ph ethers has now been demonstrated, a series of related problems, such as the rearrangement of the more important Ph ethers of mono- and polyfunctional alkanes is being investigated C J West

Constitution of oxyacanthine. II. F V BARCHHAUSEN AND P H GRICKE *Arch Pharm* 269, 113-25(1931), cf C A 24, 554 —New expl results are reported which appear to substantiate formula II for oxyacanthine and satisfactorily explain the formation of the dicarboxylic acid, 4,5-MeO(p-HO₂CC₆H₃(O)C₆H₃CO₂H) (I) therefrom The methine base of oxyacanthine Me ether previously subjected to oxidation by KMnO₄ was treated with O₃ and smoothly split according to the equation $\text{C}_{21}\text{H}_{20}\text{O}_8\text{N}_2 + 2\text{O}_3 + 2\text{H}_2\text{O} = \text{C}_{11}\text{H}_{10}\text{O}_4\text{N}_2 + \text{C}_{10}\text{H}_{10}\text{O}_4\text{N}_2$ (IV) + 2H₂O₂ into 2 cryst components, each in 95% yield The N free portion is a dialdehyde in 72° and yields a disemicarbazone and diphenyl hydrazone, it condenses readily with malonic acid to diphenyl etherdicarboxylic acid Its constitution, 4,5-MeO(p-OHCC₆H₃(O)C₆H₃CHO) (III), is easily shown through oxidation to I. The other component (IV) in 76°, is unusually sol in H₂O, possibly explained by the fact that aldehyde and carboxyl groups do not readily react with one another With MeI this product yields a cryst dimethiodide, C₂₁H₂₀O₈N₂I₂, characterized by the ease with which it is split by caustic alkali into MeCN and a cryst dialdehyde (V), 2,3,4,6-RO(MeO)₂(CH₂CH)₂C₆H₃CHO (R = 3,2,4-OHCC₆H₃(MeO)(CH₂CH)₂CO₂H-), m. 143° With mineral acids this difficultly sol product yields characteristic colors in AcOH It forms a disemicarbazone, and in CHCl₃ soln adds Br without evolution of HBr Hydrogenation with Pd-BaSO₄ leads to the formation of the satd. dialdehyde VI, which in turn yields a disemicarbazone and with hippuric acid a moroxalactone, thus showing that 1 aldehyde group is more reactive than the other In the present study the following compds were prepd and characterized 2'-methoxy-1,1'-diphenyl ether-4,5'-dialdehyde (III), m. 72° (disemicarbazone, C₂₁H₂₀O₈N₂(H₂O), m. 228°, diphenylhydrazone, yellowish white, m. 218°, oxime) 2'-Methoxy-1,1'-diphenyl ether-4,5'-dicarboxylic acid, C₂₁H₂₀O₈, m. 283°, is converted into 2'-methoxy-1,1'-diphenyl ether-4,5'-dipropionic acid, C₂₁H₂₀O₈, m. 144° Di-o-di (3-dimethylanilino-5-dimethoxydiphenyl ether-3-dialdehyde (IV) (crystalline with H₂O), m. 76° (dimethiodide, m. 259°) Di-o-methoxy-1,1'-diphenyl ether-4,5'-dialdehyde (V), m. 140°, gives with cored H₂SO₄ in AcOH soln a deep blue color (semicarbazone, C₂₁H₂₀O₈N₂, m. 220° (decomps)) Di-o-ethyl compd (VI), (by reduction



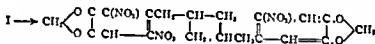
of V), m. 88-9° (disemicarbazone, m 218°; α -lactone, $C_{10}H_{11}O_2N$, yellow, m 252°)

W. O. E.

Catalytic reduction of mixtures of *p*-nitro- and nitrosophenols with aldehydes and ketones. RANDOLPH T. MAJOR. *J. Am Chem Soc* 53, 1901-8(1931)—Catalytic reduction of 2,4-(O_2N) $_2C_6H_3OH$ in Me_2CO , followed by benzoylation of the reduction product, gives 2-benzamido-4-benzoylisopropylaminophenyl benzoate, m 171°, while in 1,4-dioxane, $MeOH$, $AcOH$ or *iso*- $PrOH$ the expected product, $(H_2NH)_2C_6H_3OBz$, m 233°, is formed. 2,4-(O_2N) $_2C_6H_3OEt$ in Me_2CO gives 2,4-dibenzamidophenole, m 180°. Reduction of a mixt. of *p*- $O_2NC_6H_4OH$ and Me_2CO gives some *p*- $H_2NC_6H_4OH$, m 184° (decompn), and 4- $Me_2CHNH_2C_6H_4OH$, m 155-6° (45% yield), this was also obtained by reduction of $Me_2CNC_6H_4OH$, the *NO* deriv., yellow, m 112-3°, the phenylcarbamyl deriv m 214-5° (decompn), *HCl* salt, m 180°. Reduction of *o*- $O_2NC_6H_4OH$ in Me_2CO followed by treatment with HNO_2 gives a negligible amt. of *NO* deriv., the product being 2- $H_2NC_6H_4OH$; *p*- $O_2NC_6H_4OH$ and Et_2CO give 60% of the *NO* deriv. of 4-sec-amylaminophenol, m 89-90°; *HCl* salt, decomps without melting. Similarly $MeEtCO$ gives the 4-sec-butylamino deriv., analyzed as the *HCl* salt. When *p*- $O_2NC_6H_4OH$ is reduced in $BzMe$ or menthone, only *p*- $H_2NC_6H_4OH$ is formed. Catalytic reduction of *p*- $O_2NC_6H_4OH$ in $HClHO$ or $AcEt$ gave no definite condensation product, but with $BzEt$ in $MeOH$ a considerable yield of 4-dibenzylaminophenyl benzoate, m 174-5°. 4-Dibenzylaminophenyl benzoate, m 143-4°.

C. J. WEST

Derivatives of safrole. I. KINZO KAPUKU AND CHIUTA HATA. *J. Chem Soc Japan* 52, 20-5(1931)—A study of some derivs of 6-nitrosafrole. By oxidizing the 6-nitrosafrole (I) with $KMnO_4$, K. and H. obtained a cryst. compd m 114° which, from the reaction with the Schiff and Tollens reagents and from analysis, was considered to be 6-nitrohomopiperonal, although it formed no semicarbazone, oxime or phenylhydrazone. The other derivs studied were as follows: I \rightarrow aminosafrole \rightarrow acetylaminosafrole \rightarrow acetylaminobromopiperonylic acid \rightarrow aminobromopiperonylic acid I \rightarrow nitrohomopiperonylic acid (II) \rightarrow methylenedioxyindole II \rightarrow azoxyhomopiperonylic acid, $[HO_2CC_6H_3(CH_2O_2C)_2C_6H_3N_2O]$ I \rightarrow 5-Ac deriv. I \rightarrow 2,6-dinitrosafrole.



R. SAITO

The two stereoisomeric isoeugenols. PA. BORDACKER AND H. VOLK. *Ber* 64B, 61-6(1931)—Theoretically, aromatic propenyl compds (anethole, isosafrole, isoeugenol (I), isochavibetol) can exist in 2 stereoisomeric forms, but no representative of this type of compds had as yet been isolated in both forms in pure state. If com I, m 0-5°, is dissolved in bot 15% $NaOH$, there seps on cooling a salt which yields the I m 33° (II) and the mother liquor on further purification through a benzoate give a homogeneous liquid phenol which is not eugenol and is doubtless the stereoisomer (III) of II. II and III can be converted into each other through their dibromides. Bromination of either form gives a mixt. of the dibromides of each, although one dibromide, probably the more stable one, predominates. Thus the known acetisoeugenol, m 70°, a deriv of II, gives chiefly a dibromide (IV) m 132°, with a smaller amt. of an isomer (V) m 79°. The Ac deriv of III gives the same dibromides but the yield of V is somewhat larger. Similarly, both IV and V on debromination give mixts. of the 2 acetates. III is best converted into II by brominating its acetate, debrominating the resulting mixt., removing most of the acetate of II, again brominating the residue, etc. For the conversion of II into III the V resulting in small amt in the bromination of the acetate of II is debrominated and saponified, the II is sepd as the Na salt and the III isolated as the benzoate. By analogy, II is probably the *trans*- and III the *cis* form. II, b_p 141-2°, m 30-3°, d_4^{20} 1.0852, n_D^{20} 1.5782 (supercooled liquid) III, faintly yellow, b_p 134-5°, b_{p-1} 98°, d_4^{20} 1.0851, n_D^{20} 1.5726, benzoate, m 50-61°, depresses the m p of its isomer (m 102-3°) to 57-98° and that of eugenol benzoate (68-9°) to 45-50°. Me ether of II, m 16-7°, b_p 143-4°, d_4^{20} 1.0528, n_D^{20} 1.5692, of III, b_p 138-40°, d_4^{20} 1.0521, n_D^{20} 1.5616. The Me ether of cryst isochavibetol (m 96-7°), m 15-6-5°, b_p 147°, d_4^{20} 1.0527, n_D^{20} 1.5692, does not depress the m p of the Me ether of II, it may therefore be concluded that II and cryst isochavibetol have the same spatial configuration. Acetate of III, b_p 132°, b_p 160-2°, n_D^{20}

1.5418 Bromination of 20.6 g. of the acetate of II gives 31 g IV and 1.1 g. V; 22.8 g of the acetate of III yields 20 g. IV and 5.6 g. V.

Absorption spectra of *o*-chlorophenol-indophenol and of *o*-cresol-indophenol. MATILDA MOLDENHAUER BROOKS, *J. Am. Chem. Soc.* 53, 1820-30 (1931) — The absorption max of *o*-chlorophenol-indophenol (I) at μ_{H} 8.016 is at 625 μ . At these μ_{H} values the blue ion predominates. The absorption max of the dissolved form of *o*-cresol-indophenol (II) is at 610 μ and that of the red acid form is at 500 μ . The height of the max varies with the μ_{H} and concn of the dye. In the case of I the com sample examd was of equal purity to that of a very pure sample, while with II the com sample was only 51.54% as pure as the sample of known purity. Four curves are given.

Some relationships of the ratio of reactants to the extent of conversions of benzaldehyde and furfuraldehyde to their acetals. HOWER ADKINS, JOSEPH SEMM and LESTER M. BOLANDER, *J. Am. Chem. Soc.* 53, 1853-8 (1931) — The effect of the ratio of reactants upon the extent of the reaction of BzII with EtOII, BuOII and iso-PrOII and of furfural with EtOII has been detd when the ratio of aldehyde to alc. is varied from 1.2 to 1.10. The equil const. for the first 2 of these reactions was found to be identical for the 2 alcs but to decrease slightly with increase in ratio of alc to aldehyde. The equil const for the reaction of BzII with iso-PrOII and of furfuraldehyde (I) with EtOII appeared to be independent of the concn of reactant. The equil const. for the reaction of I and MeOII has been redctd. The following values give (1) the mols of alc per mol of aldehyde, (2) % conversion of aldehyde to acetal and (3) K : EtOII and BzII, 2.08, 25.1, 0.091, 3.00, 31.1, 0.070, 5.87, 41.9, 0.080, 7.80, 48.0, 0.082; 0.06, 53.0, 0.091. BuOII and BzII, 2.02, 24.8, 0.091, 4.17, 36.0, 0.082, 6.13, 44.0, 0.084; 8.01, 48.0, 0.080, 0.00, 52.7, 0.090. iso-PrOII and BzII, 1.83, 6.7, 0.0045, 5.86, 11.1, 0.0030, 10.41, 18.7, 0.0047. EtOII and I, 2.20, 17.1, 0.030, 6.06, 32.0, 0.030, 6.21, 32.0, 0.035, 0.80, 39.8, 0.034, 11.08, 43.7, 0.039. MeOII and I, 4.40, 30.5, 0.038. I (96 g), 270 g. dry EtOII and 0.0012 g HCl, allowed to stand 1 day, give 21% of furfural acetal, b_{D}^{20} 77-0°, b_{D}^{25} 184.5°.

Substitution in resorcinol derivatives. Bromo derivatives from β -resorcyraldehyde and their orientation. M. S. IYENGAR and H. S. JOIS, *Half-Yearly J. Mysore Univ.* 4, 199-201 (1930) — 4-Nitro-6-bromoresorcinol di-Me ether — 5,2,1-O₂N(MeO)₂C₆H₂CHO (1 g) in a slight excess of cold glacial AcOH was treated with the calcd quantity of Br in AcOH in presence of a small quantity of I as catalyst, and after standing overnight, was dild with plenty of water and filtered. On recrystn from EtOH, yellowish white crystals were obtained m 138-9° (Br 29.95%, calcd for C₁₁H₅O₂NBr, 30.5%). No oxime was obtained with the above deriv. 5,2,4-O₂N(MeO)₂C₆H₂COOH brominated in AcOH as above gave a product m 140-1°, identical with that obtained from the aldehyde (mixed m p). 4-Nitro-6-bromoresorcinol 1-Et 3-Me ether — 5,2,4-O₂N(EtO)₂(MeO)C₆H₂CHO brominated as above yielded yellowish white needles m 110-21° (Br 28.35%, calcd for C₁₁H₅O₂NBr 28.62%), identical with the brominated product of the acid. The Br deriv did not give rise to any oxime. GEORGE CALINGAERT.

Wandering of groups during brominations. M. G. SRINIVAS RAO, C. SRIKANTIA and M. SESHIA IYENGAR, *Proc 15th Indian Sci Cong* 1928, 158-9 — 5-Nitro-4-methoxy- β -resorcyraldehyde (I) on monobromination yields a normal bromonitro-aldehyde, but this on further bromination apparently rearranges to give a dibromonitro deriv, in which the aldehyde group is absent. The 2nd phenolic group in the last was ethylated and the product thus obtained was found to be identical with (1) the nitro deriv of tribromoresorcinol in which both the phenolic groups had alternatively been alkylated with MeI and EtI, resp., and (2) the dibromo (ethylated) deriv of 4-nitroresorcinol 3-Me ether, in which the nitro group is in the same position as in I. From these and a few other expts. the authors conclude that the nitro group has wandered during bromination.

Replacement of sulfoale groups by nitro groups in aromatic halogen compounds. P. S. VARMA and K. A. JOSHI, *Proc 15th Indian Sci Cong* 1928, 150, cf C. A. 14, 182.

Dihalocycanoacetyl derivatives of mesitylene. REYNOLD C. FUSON and REIO G. BEVERIDGE, *J. Am. Chem. Soc.* 53, 1985-8 (1931); cf F. and Walker, C. A. 24, 4772 — 2,4,6-Me₃C₆H₂(COCl)₂ (10 g) and 4.5 g KCN in 300 cc EtOH, boiled 4 hrs, give 50-60% of cyanoacetylmesitylene (I), m 108-9°; NaOBr gives 50-60% of the di-Br deriv. (II), Me₃C₆H₂(COBr)₂ m 81-2°, di-Cl deriv, m 38-0° (45-50% yield). I and fuming HNO₃ give 35% of the di-NO₂ deriv, m 170.3-1.3° (cor.), which yields 45% of a di-Br deriv, m 94-5°, also obtained by nitration of II. Di(cyanoacetyl)-mesitylene, m 156-7° (cor.) (70-80% yield); tetra-Cl deriv, m 52-3° (60-60% yield);

tetra Br deriv., m 121.2-22.2° (cor) (30-55% yield) All of these dihalocycanoacetyl compds have been found to be stable to cold solns of alkalis. C J WEST

New modification of the Reformatsky reaction. J A NIFUWLAND AND S FLORENTINE DALY *J Am Chem Soc* 53, 1843-6 (1931) — A β HO synthesis is obtained from α Cl esters and aldehydes or ketones and Zn, through the action of Cu metal or CuO. The catalytic action of Cu is accounted for by the formation of an intermediate between Cu and the aldehyde or ketone present, which then reacts with the Zn Cl ester compd of the Cl esters and accelerates the synthesis. C_6H_6 is the best solvent, since it brings the temp into the necessary range. It also keeps free metal surface available since the products are sol in it. Limited reaction time increases the yield. Yields are given for 10 acids in C_6H_6 , in PhMe and without solvent. α Phenyl β hydroxy- β -p-methoxyphenylbutyric acid, m 75°, Me ester, b₄ 113-5°. C J WEST

Preparation of 1,2,3-triaryl-2-hydroxybutyric acids. D IVANOV AND A SPASSOV *Bull soc chim* [41, 49, 371 5 (1931), cf *C A* 25, 2134] — PhMgBr, m and p-MeC₆H₄MgBr and p-BrC₆H₄MgBr, when added to PhCH₂CO₂MgCl and to p-ClC₆H₄CH₂CO₂MgCl condense with the formation of complexes which hydrolyze to give good yields of the corresponding triaryl substituted hydroxybutyric acids. In the course of the reaction of PhMgBr with PhCH₂CO₂MgCl, the formation of PhCH₂Br and PhCH(CO₂MgCl)MgBr (I) is assumed and these, in turn, react to give PhCH₂C(OMgBr)PhCHPhCO₂MgCl which on hydrolysis would be transformed into PhCH₂C(OH)PhCHPhCO₂H. The condensation of PhMgBr, m- and p-MeC₆H₄MgBr with PhCH₂CO₂MgCl gave 61% of 1,2,3 triphenyl 2 hydroxybutyric acid, m 176-7°, 61% of 1,3-diphenyl 2 m tolyl 2 hydroxybutyric acid, m 154-0°, and 62.5% of 1,3-diphenyl 2 p-tolyl 2 hydroxybutyric acid, m 169-70°. With p-ClC₆H₄CH₂CO₂MgCl the same Grignard reagents gave 50% of p-ClC₆H₄CH₂C(OH)PhCH(C₆H₄Cl)CO₂H, m 183°, 64.5% of p-ClC₆H₄CH₂C(OH)(m-MeC₆H₄)CH(C₆H₄Cl)CO₂H, m 157-8°, and 63.5% of p-ClC₆H₄CH₂C(OH)(p-MeC₆H₄)CH(C₆H₄Cl)CO₂H, m 175-8°. Condensation with p-BrC₆H₄MgBr yielded p-ClC₆H₄CH₂C(OH)(p-BrC₆H₄)CH(p-ClC₆H₄)CO₂H, m 181-2°. On the addn of 14.8 g of PhCH₂Br in Et₂O to I prepd from 9.65 g of PhCH₂Cl and heating for 3 hrs, 14.2 g of PhCH₂C(OH)PhCHPhCO₂H, m 176-7°, was formed. C R ADDYALL

Phenylacetaldehyde and its polymerization. JAMES R FORD *J Phys Chem* 35, 1174-9 (1931), cf *C A* 20, 1995 — PhCH₂CHO (I) on standing undergoes polymerization. The d and n increase, in general, linearly with the time, but the viscosity increases much more rapidly. Crystals of the polymer, m 104°, finally sep. The rate of formation of the polymer is approx 3 to 4 times faster at 30° than at 15°, but seems to be const. under const conditions. The solid polymer dissolves in most solvents as (C₆H₅O)₂, but in CHBr₃ as (C₆H₅O)₂. Its crystal unit is, therefore, (C₆H₅O)₂. I is not oxidized appreciably in air at ordinary temp. LOUISE KELLEY

Synthesis of 2,4,5-trimethoxyphenylalanine. T. SZÉKI AND E. LAKOS *Acta Sci Univ Francisco Josephinae, Acta Chem, Mineral, Physica* 1, 157-66 (1929) — 2,4,5-Trimethoxybenzalhydantoin was made from asarylaldehyde, and on treatment with Ba(OH)₂ and H₂O gave dl 2,4,5-trimethoxyphenylalanine, m. 217°. Its Et ester HCl, m 181°, was also produced. S S DE FENÁLY

Condensation of monochloromethyl ether with phenylmalonic ester. M GOPALA RAO AND J L SIMONSEN *Proc 15th Indian Sci Cong* 1928, 149, cf *C A* 14, 2483 — A simple method of synthesizing atropic acid is by treating PhC(CO₂Et)₂CH₂OMe (I) (from PhCH(CO₂Et)₂ and ClCH₂OMe) with 50% KOH. I with 10% alc. KOH gives methyl tropic acid, PhCH(CO₂H)CH₂OMe, b₃ 164-8° (?). No further information is given. E J C

Anacardic acid. A J HAAGEN SMIT *Proc Acad Sci Amsterdam* 34, 165-8 (1931) — EtO extrn of the fruits of *Anacardium occidentale* gave anacardic acid (I) which was purified through its Pb salt. The purified free acid crystd when seeded with pelandriac acid (II) but the mixed m. p of I and II showed some depression so it appeared that, though the two were similar, they were not identical. Hydrogenation of I in EtOAc with Pt black gave tetrahydroanacardic acid (III), m 92°. Tetrahydro-II m 98° and gave a m-p depression of 7° with III. III with FeCl₃ in EtOH gave a violet color, indicating a OH ortho to CO₂H. I lost CO₂ at 220°, the same as II, and the product (IV) b₄ 205°. The Me ether (V) of IV b₄ 202°, d₄¹⁵ 0.9240, n_D¹⁵ 1.5076, M_D 101.3 (calcd. for C₁₀H₁₄O, M_D 100.9) and HI did not split off the Me group. V gave a tetrahydro deriv., m 30°. The Me ether Me ester of I b₄ 220-2° and yielded the corresponding tetrahydro deriv. (VI), m 38°. Oxidation of VI with CrO₃ gave an acid m 59° which showed no m-p depression with palmitic acid but did give de

pressions with stearic and myristic acids, so a normal 16 C chain acid was indicated. Hydrogenation of IV gave a decabhydro deriv (VII), m 31°. VII was dehydrated with $ZnCl_2$ to give a deriv b_p 180-5°. The existence of a double bond was shown by a yellow color with $C(NO_2)_4$. This double bond was reduced catalytically and the resulting hydrocarbon, b_p 178°, n_D^{25} 1.4612, d_4^{25} 0.8323, M_D 97 (calcd for $C_{11}H_{24}$ M_D 97), was identical with pentadecylcyclohexane. The corresponding hydrocarbon from II b_p 194°, n_D^{25} 1.4506, n_D^{27} 1.4512, d_4^{25} 0.8332, d_4^{27} 0.8821, M_D 106.1 (calcd for $C_{21}H_{44}$ M_D 106.2). I is a homolog of II contg 2 less CH_2 groups and the following formula is suggested for it $C_{11}H_{22}(OH)(CO_2H)(C_{11}H_{21})$ C H PEET

2,4,5-Trimethoxybenzoic acid, a derivative of dehydrodeguelin. E P CLARK *J Am Chem Soc* 53, 2007-8 (1931). — $KMnO_4$ oxidation of dehydrodeguelin in Me_2CO gives a $HO(MeO)_2C_6H_3CO_2H$, m 210° (decompn), gives a blue color with $FeCl_3$ and gives a caronic acid (2,4,5-(MeO) $_3C_6H_2CO_2H$) on methylation. Synthesis of the acid is under way C J WEST

Condensation of ethyl phenylacetate and ethyl p -chlorophenylacetate by means of isopropylmagnesium halides. D IVANOV AND A SPASSOV *Bull soc chim* [4], 49, 375-7 (1931) cf *C A* 23, 2420. — Into a flask fitted with a reflux condenser and contg 0.15 mol of $iso-PrMgX$, 0.1 mol of $PhCH_2CO_2Et$ in Et_2O was added dropwise and the mixt was heated for 2-3 hrs. After hydrolysis and extrn with Et_2O , 25 g of Et 1,3-diphenylacetylacetate, $PhCH_2COC(Ph)CH_2CO_2Et$ m 78.0-8.5°, was formed. From $p-ClC_6H_4CH_2CO_2Et$ was similarly produced $p-ClC_6H_4CH_2COC(Ph)CH_2CO_2Et$, m 119-20°, in 93% yield. The prepn of this compd by the condensation of $p-ClC_6H_4CH_2CO_2Et$ with the aid of $NaOEt$ is not possible (*J Prakt Chem* 62, 554 (1900)) C R ADDINALL

Preparation of 1-phenyl-2,2-(dialkyl, diaryl or alkylaryl)-2-hydroxypropionic acids. D IVANOV AND A SPASSOV *Bull soc chim* [4], 49, 377-9 (1931). — Ketones react with the complex $PhCH(CO_2MgCl)MgX$ (I) to give, in good yields, substituted hydroxypropionic acids. I, prepd from 0.1 mol of $PhCH_2Cl$, was treated with 0.1 mol of the

following ketones: Pr_2CO , $MeBz$, $(PhCH_2)_2CO$, $PhBz$ and $CH_3(C(CH_3)_2)CO$, giving an 88% yield of $Pr_2C(OH)CHPhCO_2H$, m 171°, $PhMeC(OH)CHPhCO_2H$, m 182-3°, $(PhCH_2)_2C(OH)CHPhCO_2H$, m 183-4°, $PhC(OH)CHPhCO_2H$ (II), m 186-7°, and $CH_3(C(CH_3)_2)C(OH)CHPhCO_2H$, m 135°. II was also prepd from $PhBz$ and $PhCH(CO_2Na)MgBr$, this latter being readily obtained from $PhCH_2CO_2Na$ and $iso-PrMgBr$ (Cl) or $EtMgBr$ C R ADDINALL

Higher terpene compounds. XLV. Alantolactone and isalantolactone. L RUTICKA AND J A VAN MELSEN *Helv Chim Acta* 14, 397-410 (1930), cf *C A* 25, 2136. — The study of alantolactone (helenin) (I) and isalantolactone (isohelenin) (II) was undertaken because of the possible derivation of these compds from sesquiterpenes and especially to prove their relationship with santonin. I was sepd from elecampane oil together with a monocyclic sesquiterpene, b_p 135-8°, d_4^{25} 0.8864, n_D^{25} 1.5000, M_D 67.76 (calcd 67.79), which yields eudahn (III) upon dehydrogenation with Se . A liquid with a formula $C_{24}H_{40}O_2$ identical with that of I and II was also isolated.

Upon refluxing I or II (which may not be single compds) with Se , 1,7 $C_{24}H_{40}MeEt$ (IV) (styphnate m 125-6°) was obtained. IV was oxidized to 1,7-naphthalenedicarboxylic acid (0.6 g from 3 g of IV), m 294-6° ($di-Me$ ester m 86-7°) by means of $K_2Cr(CN)_6$ and KOH . Treatment of I or II with H_2 in the presence of PtO leads to the formation of tetrahydroalantolactone (V), m 143-4°, n_D^{25} 1.52°, which is either identical or stereoisomeric with deoxytetrahydrosantonin (Clemo and Haworth, *C A* 25, 708). The dihydro deriv of I, m 134°, n_D^{25} —24.6°, obtained by catalytic hydrogenation, and that of II, m 159-60°, n_D^{25} 26.2°, could not be isolated in the pure state but are definitely not identical. From the products of successive oxidation of impure I by O_3 and $KMnO_4$,

the keto acid, $CH_2=CH_2CH_2COCH_2CM_2(CH_2)_2CO_2H$, b_p 140-1°, m 92-3° (semicarbazone, m 183-5°) was isolated as the Me ester, the semicarbazone of which m 207-8°, in addn, VI, m 203-5°, was also obtained, presumably originating in a dihydroalantolactone (VII), different from either of the dihydro derivs already described. Upon reduction of V with Na and alc the glycol VIII, b_p 167-70°, was obtained (16.3 g from 31 g of V). This was transformed into the dibromide which upon treatment with quinoline at 250°, yielded the sesquiterpene IX (still contg 1% of O), b_p 130-3°, d_4^{25} 0.918, n_D^{25} 1.5104. Dehydrogenation of IX with Se gave 1-methyl-7-isopropenyl-naphthalene,

135°, which crystallizes very slowly. A compd., $b.p.$ 210°, between 1 mol of NH_4NH_2 and 2 mols of XI was also obtained by mixing the constituents with the hope of obtaining the hydrazide of X. This series of compds is to be subjected to oxidative decomposition. I. M. LEVINE

Sesquiterpene alcohol of *Cryptomeria japonica*, D. Don. Y. SUGI AND T. SENGOKU *J. Pharm. Soc. Japan* 51, 196-9 (1931), German abstr. 21—S and S isolated from *Cryptomeria japonica*, D. Don, a compd., $C_{21}H_{34}O$ (I), m 84°, $[\alpha]_D^{20}$ 19.78° (in alc). I corresponds to the *cryptomeradol* of Wienhaus and Schölz (*C. A.* 24, 1932) and did not depress the $m.p.$ of the *machilol* of Takagi. Oxidation of I gave a compd. which did not depress the $m.p.$ of *dihydroxymachilol* (m 110°), the *monobenzoate* (m 144-5°) did not depress the $m.p.$ of *dihydroxymachilol monobenzoate* (*loc. cit.*) F. I. N.

Bitter principles from alant root (preliminary communication). KARL FR. W. HANSEN, *Ber.* 64B, 67-71 (1931)—Alant root contains, in addn to the known alantolactone, $C_{11}H_{16}O_2$ (I), m 76°, and isalantolactone (II), m 115°, a 3rd bitter principle, $C_{11}H_{16}O_2$ (III), m 174°, which is also a lactone. The starting material was Merck "Helenin (alant camphor)," but as the prepn of the 3 compds in pure form by crystn. was very tedious and time consuming a process was devised which, although it involved considerable losses, gave pure, readily crystd. products, identical with those obtained by crystn. The amides of the corresponding alantolic acids, obtained with alc. NH_3 , are especially well adapted to the characterization of I and II but III gives no amide. The stability of the lactone ring in III is also shown in the behavior toward alkali, I and II yield the HO acid with boiling aq. alkalis while III requires alc. KOH. Contrary to the analyses reported by Bredt and Posth and by Sprinz for the amides obtained from I and II, which agree with the values calcd. for $C_{11}H_{16}O_2 + NH_3$, the compds obtained in this work, although having the same $m.p.s$ as those given by the earlier workers, have a quite different compn. Presumably these amides have a considerably more complicated structure than was assumed by Bredt and by Sprinz, the numerous analyses made would indicate that they contain 2 mols. lactone per mol. of NH_3 . III is apparently a dihydro deriv. of II, which gives I with Na Ilg. On dehydrogenation with Se, I, II and III give 1,7- $C_{10}H_{14}MeEt$ (IV), whose picrate (m 101°) and styphnate (m 125°) did not depress the $m.p.s$ of preps. furnished by Ruzicka who obtained the hydrocarbon by dehydrogenation of hexahydrosantenin (R. and Eichenberger, *C. A.* 25, 901) and of helenin and isohelenin (van Melsen, *Diss. Utrecht*, 1929). Of the two C atoms split off in the dehydrogenation, one presumably belongs to the C:O of the lactone ring and the other, as in eudesmol, is probably attached to a quaternary ring C atom, complete hydrogenation of the lactones and subsequent dehydrogenation with Se again gave IV and so somewhat better yield than did the non hydrogenated lactones. The 3 bitter principles are apparently sesquiterpene lactones of the eudesmol type. They were obtained by distg. the com. "Helenin (alant camphor)" in a high vacuum, treating the crystd. distillate in cold alc. with NH_3 , filtering off the alantol- and isalantolamides and recrystg. the residue from the mother liquors from AcOH. The III has a mol. wt. of 240 in camphor. The 2 amides are sepd. by crystn. from AcOH and Me_2CO and converted back into the lactones by dry distn. *in vacuo*. Amide from I, m 205-6° (Bredt and Posth, 197°), Ac. deriv., m 180°. Amide from II, m 247° (Sprinz, 237-9°), C 74.87, H 9.00, N 3.00%, Ac. deriv., m 216-7° (S, 212°), N 2.85%. IV with $K_2Fe(CN)_6$ gave 1,7- $C_{10}H_{14}(CO_2H)_2$, m 205-6° (van Melsen gives 295-6° for the pure acid) C. A. R.

Amyrin and attempts to clarify their constitution. H. DIETERLE, A. SALOMON AND E. HERZBERG *Arch. Pharm.* 262, 78-87 (1931)—The amytrins required in the present study were prepd. in accordance with a procedure duplicating in part the methods of both Ilormann and Vesterberg. By oxidation with CrO_3 , Vesterberg obtained α -amyrone, $C_{21}H_{34}O$, a result corroborated by the present authors. Entirely different results were obtained, however, when oxidation was carried out in the presence of H_2O vapor, a method previously employed by Windaus with cholesterol, and by D. with lupeol. In addn to the volatile products $AcMe$ and CO_2 , as also minute quantities of a pleasant smelling substance, a non volatile compd., $C_{21}H_{34}O$ (hydrazone, $C_{21}H_{34}N_2$, m 224°), m 89-90°, $[\alpha]_D^{20}$ 123.2°, and amyrene (hydrazone, $C_{20}H_{32}N_2$, m 252°), were isolated. While amyria is apparently unaffected by $KMnO_4$ in either alk. or acid soln., it is converted by HNO_3 into the 2 products, $C_{21}H_{32}N_2O_{10}$ and $C_{21}H_{32}N_2O_{10}$, the former m 171°, the latter, yellow, m 185-6°, both representing tribasic acids of the probable compn. $C_{20}H_{32}(NO_3)_2(CO_2H)_2$ and $C_{21}H_{32}(NO_3)_2(CO_2H)_2$. O_2 converts amyria into a dibasic and a tribasic acid, $C_{21}H_{32}O_4$ and $C_{21}H_{32}O_4$. Since these acids are unaffected by Br, they contain no double bonds. On dehydration *via* Diels, amyria

yields 3 different hydrocarbons, $C_{11}H_{18}$, $C_{11}H_{16}$ and $C_{11}H_{14}$, the constitutions of which are at present unknown. Attempts to convert α amyrin into the β form, and vice versa, were unsuccessful. On conversion of β amyrin into the formate, with subsequent sapon, an isomeric substance, γ amyrin, m 168.9°, $[\alpha]_D^{25}$ 42°, resulted, which yields the hydrocarbon γ -amyrilene, $C_{11}H_{16}$ m 175°, $[\alpha]_D^{25}$ 51.1°, reducible by H₂ to the hydrocarbon, $C_{11}H_{18}$ (amylene), thick yellow oil, b_p 120°, $[\alpha]_D^{25}$ 44.9°. W. O. F.

The Pilsch reaction. Aminocamphor and formaldehyde. H. RUFF, FR. BUXTOFF AND W. FLATT. *Helv. Chim. Acta* 13, 1026-30 (1930).—Attempts to prep dimethyl aminocamphor (I) through Pilsch's reaction (heating the amine HCl with HCHO) gave camphorquinone (II) and $(Me_2NCH_2)_2$ (III) instead of I. Aminocamphor HCl (IV) heated with H₂O gave no II and I HCl treated similarly gave only a small amt. of II. I HCl treated with HCHO gave a fair amt. of II but no III. Since the conversion of I or IV into II involves an oxidation, the following mechanism, which accounts also for the formation of III, is suggested: Me_2NH is split off from I and 2 mols. of this react with 2 mols. of $CH_2(OH)_2$ to give III, O and H₂O. C. H. PEET.

α -Carboxycamphocetan- β -acrylic and - β -propionic acids. Synthesis of β -homocamphor. F. SALMON LEGAGNEUX. *Compt. rend.* 192, 748-50 (1931), cf. Bredt, *et al.* C. A. 12, 1774.—Camphoric acid semialdehyde, $HO_2CC(CH_2)_2CHO$, treated with $SOCl_2$ then MeOH, gave the Me ester, b_p 136°. This was condensed with Zn and $BrCH_2CO_2Et$, forming the mixed ester of α -carboxycamphocetan- β -acrylic acid, $MeO_2CC(CH_2)_2CHClCOOEt$, m 53°, $[\alpha]_D^{25}$ -92.2° in MeOH. The free acid (I) obtained by sapon m 174-6°, $[\alpha]_D^{25}$ 54.8° in MeOH. I does not decolorize cold Br but can be oxidized by alk. $KMnO_4$ to camphoric acid. Reduced in presence of PtO by Adam's method it yielded α -carboxycamphocetan- β -propionic acid, $HO_2CC(CH_2)_2CH_2CO_2H$, shiny plates, m 151°, $[\alpha]_D^{25}$ 38.1° in MeOH, the Pb salt of which, heated in a current of CO_2 , gave 45% of β -homocamphor, $C_{11}H_{18}$ Cl₂ CH₂ CO, m 202-4°, $[\alpha]_D^{25}$ 111.2° in MeOH.

β -Homocamphor is very sol in org. solvents and is readily sublimed or steam-distd. Oxime, m 105°, semicarbazone, m 215°. L. E. GILSON.

Oxidation of 5-hydroxycamphor from camphorol. YASUHIKO ASAHINA AND MORIZO ISHIDA. *Ber.* 64B, 183-92 (1931).—The camphorol of Schmiedeberg and Meyer was recently shown to be a mixt. of 3- and 5-hydroxycamphor yielding on oxidation camphorquinone (or camphoric acid) and p -diketocamphane (I) (C. A. 22, 2161) and I later isolated the 5-HO compd. (II) from the camphorol (C. A. 22, 3406). By purification through the Ac deriv. or by boiling with KOH the m. p. of II has been raised to 222° and it was found that it gives I on oxidation only at high temps. and with an excess of acid, at lower temps. and in as dil. acid as possible there is formed a lower melting mixt. (around 200°) of I and a stereoisomer which decolorizes $KMnO_4$ in H₂O or Me_2CO and Br in $CHCl_3$ but on repeated crystn. is stabilized to I. The new product is autooxidizable in solid form as well as in soln. An aq. or alc. soln. which has stood a long time or a prepn. which has been exposed to the air reacts decidedly acid in H₂O or alc. The unsatd. nature of the product is undoubtedly due to enolization and it is the enol which takes up O and changes into an acid. From a prepn. a month old was isolated an acid, $C_{11}H_{16}O_4$ (III), m. 250-1°, also obtained, with considerable unchanged I, with $KMnO_4$. As III gives a disemicarbazone (m 251°), is stable toward cold $KMnO_4$, and does not reduce NH_4AgNO_3 , it is probably 1,5,5-trimethylcyclopentane 2,4-dione 1-acetic acid. 5-Acetoxycamphor semicarbazone, m 180-5°. The enol content of various preps. of the oxidation product II, as detd. by the Meyer Br titration method, ranged from 0.50% for a product m 210° to 5.70% for a product m 200° which had stood 1 month. C. A. R.

Catalytic reduction of benzidine. F. BALAS AND P. ŠEVČENKO. *Collection Czechoslov. Chem. Comm.* 3, 171-6 (1931).— $(p-NH_2C_6H_4)_2$ (I) (10 g.) in 150 cc. of glacial AcOH, shaken with 120 cc. of colloidal Pt (Skuta), contg. 2 g. of Pt, at 0.1 atm. and 50°, absorbed the theoretical amt. of H₂ for complete reduction in 18 hrs. The solvents were distd. off *in vacuo*, the residue was dissolved in abs. EtOH, the Pt was filtered off, most of the EtOH was removed *in vacuo*, the bases were pptd. with dil. NaOH, mixed with much H₂O, extd. with Et₂O and distd. at 0.8 mm. Most of the distillate was $(p-NH_2C_6H_4)_2$ (II), b_p 120°, m 59°, colorless, turning yellow brown in the air, smelting like $C_{11}H_{18}N$, strongly basic, easily sol in MeOH, EtOH and in dil. org. and inorg. acids, less sol in Et₂O, $CHCl_3$ and PhH and very slightly sol in H₂O. It does not give the characteristic reactions of I. No cryst. salts of II could be obtained but a di-HCl salt, darkening at 260° and decomp. at 290°, was obtained by passing dry HCl gas

into II in anhyd Et_2O , and a *dipicrate*, dark red needles darkening at 237° and decomg *explosively* at 247° , was obtained in abs EtOH NH_3 corresponding to over 60% of the I was obtained from the acid reduction medium and unidentified amorphous by products of high mol wt which decompd on attempted distn were left after extrn of the II. *Bicyclohexane* was not found C H PEET

Addition of free radicals to unsaturated compounds (preliminary paper). J B CONANT AND H W SCHIFFR *J Am Chem Soc* 53, 1941-4 (1931)— Ph_3C and $\text{CH}_2=\text{CMeCH Cl}_2$ in N or Ph_2CCl_2 , $\text{CH}_2=\text{CMeCH Cl}_2$ and Hg in C_6H_6 give 1,4 *bis*(triphenylmethyl) 2 methyl 2 butene, m $163-4^\circ$, the structure was established by the action of O_2 , giving $\text{Ph}_3\text{CCH}_2\text{CO}_2\text{H}$ and 1,1,1 triphenylbutan 3 one, m $140.5-1^\circ$, also obtained from $\text{Ph}_3\text{CCH}_2\text{COCl}$ and MeZnCl $\text{CH}_2=\text{CMeCMe CH}_2$ and Ph_3C give the compd $\text{C}_{41}\text{H}_{48}$, m 240° (decompn). O_2 gives $\text{Ph}_3\text{CCH}_2\text{Ac}$ C J WEST

Action of cyanoacetic acid on triphenylcarbolol. Synthesis of triphenylmethylcyanoacetic and *N*-triphenylmethylmalonamic acids. R FOSSE *Bull soc chim* [4], 49, 159-73 (1931) cf C A 1, 2797— Ph_3COH reacts with $\text{NCCH}_2\text{CO}_2\text{H}$ to give $\text{Ph}_3\text{CCH}(\text{CN})\text{CO}_2\text{H}$ (I) by elimination of H_2O and also by addn, forming a new type of acid, *N*-triphenylmethylmalonamic acid, $\text{Ph}_3\text{CNHCOCH}_2\text{CO}_2\text{H}$ (II), the λ Ph_3C deriv of the yet unknown malonamic acid. The sirup like liquid formed by heating 10 g of Ph_3COH in 25 g of $\text{NCCH}_2\text{CO}_2\text{H}$ for 4 min on the steam bath was freed from excess $\text{NCCH}_2\text{CO}_2\text{H}$ by treatment with H_2O , and was agitated with 100 cc anhyd Et_2O in a stoppered flask. On addn of petroleum ether to the ether soln dried over Na_2SO_4 , and on standing for 24 hrs crystals were formed which gave off H_2O at 60° in *vacuo* to form I, m 155° . Heated to $160-70^\circ$, CO_2 was evolved with formation of $\text{Ph}_3\text{CCH}_2\text{CN}$ (III), m 140° after recrystn from EtOH (C A 23, 4025) III was also prepd by the action of Ph_3COH on $\text{NCCH}_2\text{CO}_2\text{H}$ at $110-5^\circ$, and by evapn to dryness of a soln of $\text{Ph}_3\text{CCH}(\text{CN})\text{CO}_2\text{Na}$. To establish the constitution of I and III, the latter was hydrolyzed by heating with alc KOH to 280° for 20 min to $\text{Ph}_3\text{CCH}_2\text{CO}_2\text{H}$, which in turn was prepd by the fusion of Ph_3COH with $\text{CH}_3(\text{CO}_2\text{H})$. $\text{Ph}_3\text{CCH}_2\text{CONH}_2$ was formed by refluxing 10 g of $\text{Ph}_3\text{CCH}_2\text{CO}_2\text{H}$ with 3.5 g of PCl_5 in 60 cc dry PhH for 20 hrs, and passing dry NH_3 into the cooled soln (C A 21, 2670). The amide was then transformed into III by refluxing with PCl_5 in PhH for 4 hrs. The ether insol residue (II) from the prepn of I, when recrystd from CHCl_3 , m 175° , giving off CO_2 and forming Ph_3CNHAc (IV), m 211° , identical with the product of the action of Ac_2O on Ph_3CNH_2 (Ber 17, 744 (1884)) or by the fusion of Ph_3COH with AcNH_2 at $210-40^\circ$ for 45 min in the presence of a trace of H_2SO_4 . It was hydrolyzed by 90% H_2SO_4 to Ph_3COH , NH_3 and AcOH . II is therefore $\text{Ph}_3\text{CNHCOCH}_2\text{CO}_2\text{H}$, m 175° C R ADDINALL

Combinations among certain dyestuff radicals. M REBEK. *Collection Czechoslov Chem Comm* 3, 155-70 (1931)—(*p*- $\text{Me}_2\text{NC}_6\text{H}_4$) COH (I) condensed with [2,4-(O_2N) $_2$ - C_6H_3] CH in boiling Me_2CO in PhNO_2 at $120-5^\circ$ or by melting the components together, to give a compd (II) which was pptd by addn of Et_2O . On recrystn from Me_2CO , II yielded large green crystals with a golden glance which analyzed as $\text{C}_{24}\text{H}_{15}\text{O}_8\text{N}_5$, $\cdot 3\text{Me}_2\text{CO}$. Mol-wt detns (i p) in PhNO_2 indicated 2 mols of Me_2CO of crystn. Similarly, I condensed with [2,4,6-(O_2N) $_3$ - C_6H_2] NH (III) to give long green crystals which analyzed as $\text{C}_{24}\text{H}_{15}\text{O}_8\text{N}_6$, $\cdot \text{Me}_2\text{CO}$ (IV) and this formula was confirmed by a mol-wt. detn. IV reacted with I, in CHCl_3 to give a trioxide, 3,4 $\text{Me}(\text{H}_3\text{N})\text{C}_6\text{H}_3\text{C}(\text{OH})(\text{C}_6\text{H}_4\text{N}_2\text{H}_3(p))_2$, condensed with III to give $\text{C}_{24}\text{H}_{15}\text{O}_8\text{N}_8$ (V). The rate of reaction for each of these condensations was followed via the cond changes of their Me_2CO or PhNO_2 solns and these data are tabulated. The reactions which give IV and V were instantaneous but, probably because both the pseudo base and pseudo acid involved must change to the true base and true acid form preliminary to the formation of the dye salt, II was formed slowly. Color changes occurred at definite dilns. II and IV exhibited identical absorption in the yellow and green regions and spectral examn indicated that the optical effect is additive. An approximation to Lambert and Beer's law was observed C H PEET

Triphenylmethane dyes derived from quinoline, tetrahydroquinoline, diphenylamine and carbazole. RAJENDRA NATH SEN AND BHOLA NATH SEN *J Indian Chem Soc* 7, 965-72 (1930)—Two mols of quinoline (I) condense with 1 mol of an aromatic aldehyde in the presence of concd HCl to form a leuco base oxidizable by PbO_2 to a carbinol, the linkage taking place at the 6-position of I since the 6-nitro deriv. of I did not condense. The reaction is greatly retarded by the presence of a *m*-acidic substituent in the aldehyde. Aliphatic aldehydes did not condense. Tetrahydroquinoline (II) and its *N*-Me deriv condensed more easily than I, the greater activity of the 6-H and stronger basicity of the products being due to the ring N atom which is doubly linked

in I, but singly in II to form a secondary (alkylated) amine. Methylation of the amino group as in Ph₂NH greatly retards the reactivity of the *p*-II, but methylation of the amino II in this compd. markedly increases the reactivity of the *p*-II. Carbazole condensed less readily and methylation had less effect, showing that the influence of the NH group is affected when the N forms part of a ring. The II complex exerts a stronger auxochromic influence than the I complex. The following diquinolylmethanes were prepd: *phenyl*, by heating 14 cc I, 6 cc BzII, and 15 cc concd. HCl 26 hrs. on the water bath, extg with 100 cc H₂O, making alk with NaOH, removing excess BzII by steam distn., repptg. the product from AcOH with NH₄OH, and crystg. from equal parts of EtOH and AcOH (yield 30%), dyes wool and silk light green (chloroplatinate, m 230°, picrate, m 199°). *p*-methoxyphenyl, from I and anisaldehyde, orange-yellow; *o*-hydroxyphenyl, m 157-60°, light yellow, *p*-hydroxyphenyl, m above 263°, brownish red, *p*-nitrophenyl, m 108°, light green (the *o*- and *m*-comps were not obtained pure), dimethyl-*p*-aminophenyl, m 168-70°, brown, 2,4-dihydroxyphenyl, m. above 240°, orange (*Ac deriv.*, decomps without melting; *di Ac deriv.*, m above 267°, *di Ag salt*, red flakes). *Ditetrahydroquinolylmethanes*: *m*-hydroxyphenyl, by heating 6 g II and 2.8 g m HOC₆H₄CHO, with 25 cc. 20% HCl 5 hrs. on the water bath, and working up as before, m above 256°, light brown, *m*-nitrophenyl, m 212° (decompn.), bright green, *furfuryl*, *p*-methoxyphenyl, light bluish green, *o*-hydroxyphenyl, m 163-4°, bluish green, *p*-hydroxyphenyl, m 225°, bluish ash; dimethyl-*p*-aminophenyl, softens about 10°, yellowish brown, 2,4-dihydroxyphenyl, m above 235°, reddish brown. *p',p'*-Diphenyldiaminotriphenylmethane, by refluxing for 13 hrs 10 g Ph₂NH, 3 cc. BzII, and 30 cc. EtOH, and recrystg. the insol product from EtOH and AcOH (1.2), green powder (*disulfonic acid*, by heating 5 g with 15 cc. concd. H₂SO₄ 10 hrs on the water bath, green plates from EtOH), *o*-HO *deriv.*, from Ph₂NH and salicylaldehyde, green powder, 3,4-methylenedioxy *deriv.*, from Ph₂NH and piperonal, green powder. *Diphenyldiaminodiphenylmethane*, from Ph₂NH and HCHO, bluish green powder. *Dimethyldiphenyldiaminotriphenylmethane*, by heating Ph₂NMe and BzII 0.5 hr on the water bath without a condensing agent, m 144-6°, greenish. *Dimethyldiphenyldiaminodiphenylmethane*, from Ph₂NMe and HCHO, dark green powder. *Phenyldicarbazolemethane*, from carbazole and BzII, m. above 260°, faint green, *di N-Me deriv.*, from *N*-methylcarbazole and BzII, m 149°, bluish green. R. C. ELKRAFT.

Mixed benzoin. IV. Determination of the structure of mixed benzoin by the Beckmann reaction. JOHANNES S. BUCK AND WALTER S. INA. *J. Am. Chem. Soc.* 53, 1912-7 (1931), cf *C. A.* 25, 2713.—The structures of 7 mixed benzoin, obtained by the usual KCN condensation have been detd. by means of a 2nd type Beckmann reaction. The results, interpreted according to Werner's views of the Beckmann rearrangement, give structures in agreement with those detd. by independent methods. The compounds studied were C₁₂H₁₁CH(OH)COC₆H₄(OMe)₂, C₁₂H₁₁CH(OH)COC₆H₄OMe, C₁₂H₁₁CH(OH)COC₆H₄NMe₂, C₁₂H₁₁CH(OH)COC₆H₄OCH₃, PhCH(OH)COC₆H₄OMe, PhCH(OH)COC₆H₄O₂CH₃, PhCH(OH)COC₆H₄NMe₂. The oximes of these m as follows 149°, 144°, 158-8°, 149°, 140°, 158°, 140°. The yields of aldehyde and nitrile obtained in the Beckmann reactions are listed.

C. J. WEST.
Cyclohexanone-benzil. CHARLES F. H. ALLEN. *Can. J. Research* 4, 264-74 (1931).—Cyclohexanone and benzil in the presence of a small amt. of NaOMe formed an addn. product which is a "semicyclic" 1,4-diketone. No derivs. of a bicyclic cycloheptane were found. Its most conspicuous chem. property is its sensitivity to mineral acids. Among the reaction products was found a diphenyltetrahydrocoumarone. Although the addn. product exhibits the phenomenon of halochromism the existence of a perchlorate is doubtful and it did not form a free radical. Analysis of cyclohexanone-benzil gave C 77.5, H 6.5%, mol wt 300-6 (calcd for C₂₁H₂₂O₂: C 77.9, H 6.5%, mol wt. 308). Reactions of cyclohexanone-benzil with the OH and C O groups and with acidic and dehydrating reagents were studied.

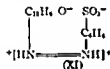
J. W. SHIPLEY.
Naphthalene series. III. Action of bisulfite on *p*-nitrobenzenecarbo- β -naphthol. N. N. VOROZHTZOV AND P. A. BIELOV. *Ber.* 64B, 77-83 (1931), cf *C. A.* 23, 2434, Turski, Pulawski, Hildebrand and Bortnowoska, *C. A.* 23, 2037.—Acid sulfites with azo dyes generally give bisulfite compds. from which the original dye can be regenerated by hydrolysis or by adding alk. reagents to their aq. soln. Of these dyes *p*-nitrobenzenecarbo- β -naphthol (pared) (I) was of special interest because of its practical importance and also of the presence of a NO₂ group, which is not indifferent toward SO₂ and sulfites. V. had found 15 yrs. ago that I, probably because of its slight soly., reacts with NaHSO₃ considerably more slowly than the other azonaphthol dyes studied, yielding in every case as an end product the yellow crystals of a bisulfite compd. (or a mixt. of several), if the reaction is carried out under sufficiently rigorous conditions.

the bisulfite compd. on decompn. yields a dye quite different from the original I; it contains S (apparently as SO_3H), dissolves in aq. alkalis (even carbonates) and contains a primary diazotizable NH_2 group. The new dye gives with excess of NaHSO_3 the bisulfite compd. to be expected of naphthol dyes. These expts. are described in a sealed report filed in 1916 with the *Gesellschaft zur Mitwirkung an der Verbesserung und Entwicklung der Textilindustrie in Moskau*, and the work has been taken up again because of the appearance of the paper by Turski, *et al.*, who found $\text{H}_2\text{NC}_6\text{H}_4\text{N} \cdot \text{NC}_6\text{H}_4\text{OH}$ and $\text{HO}_2\text{S}(\text{H}_2\text{N})\text{C}_6\text{H}_4\text{N} \cdot \text{NC}_6\text{H}_4\text{OH}$ among the products. As far as can be judged from the *Chem. Zentr.* abstr. of their paper, they give no complete explanation of the course of the reaction, and A and B have endeavored especially to det. whether the reaction begins with the reduction of the NO_2 group or whether the formation of the bisulfite compd. is the 1st step. They found that under mild conditions I and NaHSO_3 give a compd. (II) of the compn. $\text{LN}(\text{NaHSO}_3)$ which readily regenerates the original I and has all the properties (as regards hydrolytic cleavage by diln., by acids or by alkalis) characteristic of the azonaphthol dye derivs. of this type. In agreement with the views previously developed, II must have the structure $\text{C}_6\text{H}_4 \cdot \text{CH} \cdot \text{CH} \cdot \text{C}(\text{OH})(\text{OSO}_3\text{Na}) \cdot \text{C} \cdot \text{N} \cdot \text{NHC}_6\text{H}_4\text{NO}_2$. The II is obtained in 25 g. yield

by refluxing 14.65 g. I with 75 cc. CHCl_3 , 150 cc. alc., 125 cc. NaHSO_3 soln. (33.8% NaHSO_3) and 70 cc. H_2O 20–5 hrs., filtering hot from the unchanged I and cooling to 25° . It is sol. in about 150 parts H_2O , forming orange-red coned. and yellow dil. solns., in coned. aq. soln. it begins to decomp. at $60\text{--}5^\circ$. With coned. solns. of CaCl_2 , BaCl_2 , NiCl_2 , CuSO_4 , $\text{Pb}(\text{NO}_3)_2$, AgNO_3 , it at once forms ppts. which are yellow in the presence of H_2O but assume a deeper color when dried in the air or in desiccators, the *Ba* compd. has the compn. $\text{C}_{22}\text{H}_{13}\text{O}_{12}\text{N}_2\text{S}_2\text{Ba} \cdot \text{H}_2\text{O}$, and the *Cu* salt contains $8\text{H}_2\text{O}$. The dissocn. of II in aq. solns. of various concns. at 25° was followed by iodometric titration, the amt. of titratable SO_2 in % of total SO_2 after 24 hrs. increased from 6.68 in $1/100$ to 11.08 in $1/100$ N soln. In 50 cc. 0.01 N soln. of II, 25 cc. H_2O and varying amts. of NaHCO_3 , the titratable SO_2 increased from 5.58% with 0 equiv. NaHCO_3 to 73% with 10 equivs. NaHCO_3 . In 25 cc. 0.01 N II and 25 cc. H_2O with varying amts. of HCl , the titratable SO_2 decreased from 9.2% with 0 equiv. HCl to 3.84% with 0.03 equiv. HCl and 1.56% with 10 equivs. HCl . C. A. R.

Constitution of the azo indicators. I. α -Naphthol orange. K. H. SLOTTA AND W. FRANKS. *Ber.* 64B, 86–94 (1931).— α -Naphthol orange (I), compared with other compds. of similar mol. wt. contg. a SO_3H group, is strikingly little sol. in H_2O , indicating that there is some point in its mol. which can extremely readily bind the H nucleus of the SO_3H acid to form an inner salt. This was confirmed in attempts to det. the acid const. of I by cond. methods, the cond. of I in H_2O is so small (of the order of magnitude of the cond. water used) that it could not be detd. The question now again arose whether the dissocn. of one H atom suffices for the indicator color change or whether the H of the phenolic OH must also dissoc. Conductometric titration of I showed that the color change to red occurred simultaneously with the 1st break in the titration curve on addn. of exactly 1 equiv. of alkali; the dissocn. of the 2nd H on further addn. of alkali, manifested by a 2nd break in the curve when the 2nd equiv. of alkali has been added, is without optical effect. This 2nd H atom can smoothly be replaced by alkyls in alk. soln. and the alkylated products do not lose their indicator properties. Alk. Me_2SO_4 gives only the mono-Me deriv. (II), almost identical with I itself as regards the color change interval and the half-value stage (i. e., the pH at which 50% of the red and yellow forms are present). The same is true of the other monoalkyl derivs. obtained with Et_2SO_4 , $\text{p-MeC}_6\text{H}_4\text{SO}_3\text{Pr}$, $\text{p-MeC}_6\text{H}_4\text{SO}_3\text{Bu}$ and *hexyl* bromide, the introduction of the alkyls merely deepening the limiting colors. The covering power is also materially increased. Below are given the properties of I and its *Me* (II), *Et* (III), *Pr* (IV), *Bu* (V) and *hexyl* derivs. (VI), resp.: Color change interval, 7.4–8.9, 7.4–9.0, 7.3–8.9, 7.4–8.9, 7.3–9.0, 7.4–8.9; half-value stage, 8.27, 8.25, 8.23, 8.26, 8.33, 8.27; alk. color, gray pink, blue-pink, rose-crimson, crimson, crimson, crimson; acid color, pale yellow green, yellow green, yellow, golden yellow, golden yellow, golden yellow, amt. of indicator for equal covering power, 100, 90, 75, 70, 75, 60. In view of this similarity in properties, all these indicators must have the same structure. The Me group in II is not split off by const.-boiling HI , indicating that it is not on the phenolic HO or the SO_3H group. Only with NH_4I above 200° is MeI split off, and even then only 0.5 of the calcd. amt. is given off. $\text{p-H}_2\text{NNMeC}_6\text{H}_4\text{SO}_3\text{H}$ (VII) behaves in exactly the same way, showing that in II the Me is on one of the azo N atoms. With alk. $\text{Na}_2\text{S}_2\text{O}_4$ II gives as one of the cleavage products 1,4- $\text{C}_6\text{H}_4(\text{OH})\text{NH}_2$ (VIII) which was detected by covering the reaction mixt. with Et_2O ; the VIII

dissolved in the 1:1:1 with blue violet color as fast as formed and, after evapn of the ether, was isolated as the yellow *p*-nitrobenzoyl deriv, decomps 276°, by treatment in $C_{12}H_5N$ with $O_2NC_6H_4COCl$. This shows that the Me in II is on the N attached to the $C_{12}H_5$ ring and hence that it is not the II on this N which is responsible for the indicator properties of these compds. The other N atom could not be alkylated with any of the above agents but solid II in 1:1:1 suspension gave on long standing with $C_{12}H_5N$, a red orange di Me deriv (IX) which is insol in H_2O and in an Me_2CO and shows no color change with alkalis, with $Na_2S_2O_4$ in hot $MeOH$ it yields 1,4 $C_{12}H_4(OH)NHMe$ (X) which splits off no MeI with $HI-P$ and is distinctly sol in alkalis, showing that the Me is on the N and not on the HO group. These facts indicate that the yellow 'acid' form of I has the double zwitterion structure XI; as the pH of the soln is raised a H nucleus is withdrawn from the azo N atom whose attached ring is best adapted to a neutralization of the union set free by becoming quinonil (the $C_{12}H_5$ nucleus), giving the deeply colored *p*-naphthoquinonephenylhydrazonesulfonic anion (XII), whose hydrazone II can be substituted by the usual alkylating agents. If these views are correct, it should be possible to obtain the di Me deriv directly from the 'acid' form of I, and as a matter of fact the solid I allowed to stand some hrs with $C_{12}H_5N$ in Et_2O changes into the red di Me deriv I (90% from diazonizenesulfonic acid and α naphthol in HCl/HOH), decomps 210°. II (yield, almost quant.), needles with green luster, decomps 210°. *p*-Methoxyazobenzene *p*-sulfonic acid, which is the product obtained under the same conditions from $HOC_6H_4N_2NC_6H_4SO_3H$, is light brown and gives no color change with alkali. III (36 g from 33 g I), green brown, decomps 263-5°. IV (75% yield), green black, decomps 268°. V (64%), decomps 269-70°. VI (22%), green, decomps 267°. IX, m 178°. *N*-*p*-nitrobenzoyl deriv of X, yellow, m 210°.



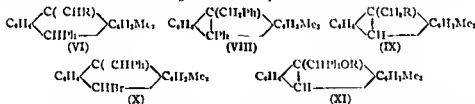
C. A. R.

Asymmetric syntheses. IX. Preparation of optically active substituted glycolic acids from *l*-menthyl and *l*-bornyl α -naphthoylformates. ALEX. MCKENZIE and PATRICK D. RITCHIE. *Biochem Z* 231, 412-22 (1931). — *l*-Menthyl α -naphthoylformate, α $C_{12}H_5COCO_2C_{12}H_5$ (I), m 88.5-9°, was prepd from naphthoyl formamide (II) and *l*-menthol. It showed slight mutarotation in $PrOH$, $iso-BuOH$, and $EtOH$, the l rotation increasing. $PhMgBr$ and I on hydrolysis give phenyl α -naphthylglycolic acid, both as racemic and optically active acid, $[\alpha]_D^{25} -16.4^\circ$. Similarly, methyl and ethyl α -naphthylglycolic acids were obtained as *dl* and *l* acids, $[\alpha]_D^{25} -30.7^\circ$ (methyl), $[\alpha]_D^{25} -6.8^\circ$ (ethyl). *l*-Bornyl α -naphthoylformate (III), m 69.5-70.5°, was prepd from II and *l*-borneol and showed mutarotation in $EtOH$, $[\alpha]_D^{25} -26^\circ$ to -27° . Methyl α -naphthylglycolic acid prepd from III and $MeMgI$ occurred as the *dl* and *l* acids, $[\alpha]_D^{25} -34.1^\circ$. D. S. SEARLE

Anthracene series. Halogenation and nitration. P. S. VARMA AND A. SUBRAMANYAM. *Proc 15th Indian Sci Cong* 1928, 151. — Monoiodo- and diiodo-anthraquinone were obtained by the action of $NaNO_2$ and fuming H_2SO_4 on anthracene and I. A number of Br compds were obtained by the action of Br and fuming H_2SO_4 . Some of these halogenated compds were also nitrated. In some cases nitro derivs were obtained, while in others the halogen compds were decompd and the halogens liberated. No details are given. E. J. C.

Anthracene derivatives. III. EDWARD DE BARRY BARNETT AND JOHN A. LOW. *Ber* 64B, 49-56 (1931), cf *C A* 24, 1638. — The work of the last 10 yrs has agreed with the hypothesis that the influence of a Cl atom in the reactions of certain anthracene derivs is due to a coordination with a H atom in the *peri meso* position but it was desirable to show that the properties produced by a coordinated group would not be produced by a non-coordinated group in the same position. An investigation of the α -methylanthracenes was accordingly undertaken and, contrary to expectations, it was found with 1,4-dimethylanthracene (I) that an α -Me group has very much the same influence as an α -Cl atom. 1,4-Dimethyl 10-bromoanthrone (II) readily yields the corresponding 10-MeO deriv (III) and reacts with bases, $NHRR'$, to form the 10-NRR' derivs (IV). II also readily undergoes the Friedel-Crafts reaction with C_6H_6 , the resulting 10-Ph deriv (V) giving with Grignard reagents carbinols (not isolated) which, under the catalytic influence of H ions, undergo no transannular de-

hydration but change into alkylidene compds (VI), whose structure was proved by their complete lack of color and the absence of fluorescence, and also by the oxidation of the methylene deriv ($R = H$) back to V. That the intermediate carbinols, however, can readily undergo transannular dehydration when alkylidene formation is impossible is shown by the very smooth formation of 1,4 dimethyl 9,10 diphenylanthracene (VII) from V with $PhMgBr$ and by the formation of I (see below). Furthermore, the benzylidene deriv VI ($R = Ph$) boiled with $AcOH$ contg HCl gives a yellow, strongly fluorescent substance, probably VIII. 1,4 Dimethylanthrone with Grignard reagents and subsequent treatment with HCl gives the faintly yellow, strongly fluorescent compd IX; the benzyl compd ($R = Ph$) gives a Br deriv (X) which forms non fluorescent $NR'R''$ and $R'O$ derivs with bases $NHR'R''$ and $als. R'OH$, resp. The latter on boiling with als contg a little HCl change into easily sol, yellow, fluorescent substances, probably of the structure XI. Although v Brann and Bayer were unable to acetylate 1,4 dimethylanthrone with hot Ac_2O $NaOAc$ the 1,4 dimethylanthranyl acetate (XII) was easily obtained with C_6H_5N Ac_2O . V, however, is recovered unchanged after heating 2 hrs with C_6H_5N Ac_2O on the water bath. With Zn dust and NH_4OH on the water bath, 1,4 dimethylanthrone gives the 9,10 dihydroanthranol which on catalytic dehydrogenation in the presence of H ions readily yields I. I at once adds Br but the resulting dibromide slowly loses HBr at room temp and both HBr and Br in boiling C_6H_6 . It is clear therefore, that in spite of the impossibility of a *peri* coordination the derivs of I are much more similar to the corresponding derivs of 1,4-, 1,5- and 4,5-dichloro and 1,5 diphenoxanthracene than to those of anthracene itself, and it must be concluded that the hypothesis of *peri* coordination is untenable. It remains to be seen how many of the explanations on an electronic basis which have in recent yrs been advanced by some investigators will stand the test of a more thorough and critical examn than they have yet received. An extensive study of the influence of *bs*-substituents on the *meso* position of the anthracene complex has been begun and in the meantime no further theoretical discussion will be indulged in. α -(2,5-Me $_2$ C $_6$ H $_3$ CO) C_6H_4 CO $_2$ H, from $AlCl_3$ slowly added to $C_6H_5(CO)O$ in p -xylene, m 148° , reduced by Zn dust (activated with $CuSO_4$), NH_4OH and $NaOH$ to 2,5 dimethylidiphenylmethane 2' carboxylic acid, m 136° , which with 80% H_2SO_4 on the water bath gives almost quantitatively 1,4 dimethylanthrone, m 116° . XII, m 162° , forms yellow, fluorescent solns. 1,4,9 Trimethylanthracene (IX, $R = H$) m 81° . 4 benzyl compd ($R = Ph$), m 135° . 1,4-Dimethyl 10 bromo 9 benzylidene 9,10 dihydroanthracene (X), very light yellow, m 150° , 10 piperidino compd, m 150° , 10 MeO compd, m $118-9^\circ$, 10 EtO compd, m 136° . II, from the anthrone with Br in CS_2 , yellow, decomps about 160° , 10 piperidino compd (IV), m 133° , 10- $PhNH$ deriv, bright yellow, m 192° , 10 $PhNMe$ deriv, yellow, m 179° , 10 *p*-dimethylaminophenyl compd, m 154° , III, m 76° , V, m 141° . 1,4-Dimethyl 9 methylene 10 phenyl 9,10 dihydroanthracene (VI, $R = H$), from V and $MeMgI$, m 129° , 9 benzylidene compd (VI, $R = Ph$), from V and $PhCH_2MgCl$, m 170° . VII, light yellow, m 189° , shows vivid blue fluorescence in soln in the light of an arc lamp.



C. A. R.

ms-Dichloroanthracene and its β -sulfonic acid as starting materials for production of alizarin. V. I. MINAEV AND B. P. FEDOROV. *Rev. gén. mat. color* 34, 330-2, 376-82 (1930).—See C. A. 25, 1252.

I. J. C.

3-Nitrophthalic acid series. II. W. UNDERWOOD, JR., AND R. L. WAKEMAN. *J. Am. Chem. Soc.* 53, 1839-42 (1931).—Condensation of 15 g. 3-nitrophthalic anhydride (I) and 1 g. $PhOH$ with 3 g. concd H_2SO_4 by heating 2.5 hrs at 170° and 3 hrs at $215-20^\circ$ gives 2.1% of phenol-3-nitrophthalate, yellow, m $204-5^\circ$, the $NaOH$ soln has a violet color, a very dil soln of the Na salt becomes practically colorless on addn of a large excess of concd $NaOH$. I (5 g.) and 6.3 g. m $C_6H_4(OH)_2$, heated 1.25 hrs at $105-200^\circ$, give 17.4% of resorcinol 3-nitrophthalate, red, m 260° , the $NaOH$ soln. is brownish yellow; a dil alk soln exhibits a green fluorescence but not nearly to the same extent as a soln of fluorescein. The 1-mono-Me 3-nitrophthalate is un-

changed on heating with Ac_2O , while the 2 mono Me ester gives the anhydride, the di Me ester is unchanged. The isomeric 3 nitrophthalamic acids and 3 nitrophthalamide are converted into 3 nitrophthalimide by treatment with Ac_2O . C. J. W.

A reaction of aromatic 1,4-diketone monoximes occurring instead of the Hofmann rearrangement. I. ROLAND SCHOLL, HANS SEMPE and EDMUND STIX. *Ber* 64B, 71-7 (1931).—The monoxime of α (2,4 $\text{Me}_2\text{C}_6\text{H}_3(\text{CO})_2\text{C}_6\text{H}_4$), with alc. HCl at 150° gives the xylidide $\text{Me}_2\text{C}_6\text{H}_3\text{NHCOC}_6\text{H}_4(\text{CO})_2\text{C}_6\text{H}_4$ (I) and was therefore assigned the structure $\text{Me}_2\text{C}_6\text{H}_3\text{C}(\text{NOH})\text{C}_6\text{H}_4(\text{CO})_2\text{C}_6\text{H}_4$ (II) (*C. A.* 21, 2084). The yield of I was very small, however, and attempts were made to increase it by using other rearranging agents than alc. HCl . $\text{AcOH}\cdot\text{H}_2\text{SO}_4$ gave in good yield a yellow, weakly basic compd $\text{C}_{22}\text{H}_{19}\text{O}_3\text{N}$ (III), contg 1 H_2O less than I. Of the 2 theoretically possible Beckmann rearrangement products of II, viz. I and α (2,4 $\text{Me}_2\text{C}_6\text{H}_3\text{CONH})\text{C}_6\text{H}_4(\text{CO})_2\text{C}_6\text{H}_4$ (IV), neither gives III with $\text{AcOH}\cdot\text{H}_2\text{SO}_4$. I splits off $\text{C}_6\text{H}_4(\text{CO})_2\text{C}_6\text{H}_4\text{CO}_2\text{H}$ and IV gives $\text{C}_6\text{H}_4(\text{CO})_2\text{C}_6\text{H}_4\text{NH}_2$. The formation of III, therefore, is not dependent on a Beckmann rearrangement. The most plausible assumption, then, was that anhydridization occurred between the oxime OH and a H of the 2 Me group in the $\text{Me}_2\text{C}_6\text{H}_3$ residue, with formation of the isoindole, $\text{C}_6\text{H}_3\text{C}_2\text{H}_2\text{Me CR N}$ ($\text{R} = \text{C}_6\text{H}_4$,

$\text{CO})_2\text{C}_6\text{H}_4$), especially as the oxime of *p*-xylol α -anthraquinonyl ketone (V) gave with $\text{AcOH}\cdot\text{H}_2\text{SO}_4$ an "anhydride" (VI) with the same ease as does II. But the oxime of the *p*-tolyl (VII) and Ph homologs (VIII) also give "anhydrides" with great ease, while 2,4 $\text{Me}_2\text{C}_6\text{H}_3\text{C}(\text{NOH})\text{Ph}$ forms no anhydride with either $\text{AcOH}\cdot\text{H}_2\text{SO}_4$ or any other dehydrating agent but merely undergoes the Beckmann rearrangement. The H_2O split off in the formation of III can therefore be formed only from the oxime OH and aromatic H. To det. at what point the anhydride ring was closed the oximes of 2,1- $\text{C}_6\text{H}_4(\text{CO})_2\text{C}_6\text{H}_3\text{MeCOR}$ ($\text{R} = \text{Ph}$ and 2,4- and 2,5- $\text{Me}_2\text{C}_6\text{H}_3$) and of α (2,4,6 $\text{Me}_3\text{C}_6\text{H}_2\text{CO})\text{C}_6\text{H}_3(\text{CO})_2\text{C}_6\text{H}_4$ were treated with $\text{AcOH}\cdot\text{H}_2\text{SO}_4$. None formed an anhydride, the first 3 were unchanged by short treatment and under more drastic conditions gave black, indefinite products together with a little 2,1 $\text{C}_6\text{H}_4(\text{CO})_2\text{C}_6\text{H}_3\text{MeCO}_2\text{H}$, while the last reacted in some other way which remains to be elucidated. These results

led to testing whether the anhydrides are phthalylacridines, $\text{C}_6\text{H}_4(\text{CO})_2\text{C}_6\text{H}_3\text{C}_2\text{H}_2\text{N} \begin{smallmatrix} \text{CH} \\ \diagup \diagdown \\ \text{N} \end{smallmatrix} \text{C}_6\text{H}_4$.

formed directly from the oximes or from their tautomeric nitroso forms $\text{C}_6\text{H}_3(\text{CO})_2\text{C}_6\text{H}_3\text{CH}(\text{NO})\text{Ph}$. Phthalylacridines of this type are not known, but the corresponding acridones are. Both Ullmann's anthraquinone 1,2 acridone and the "anhydride" obtained from VII were distd over Zn dust, the former yielded β -aminooanthracene (IX), also obtained by distn of β $\text{C}_6\text{H}_4(\text{CO})_2\text{C}_6\text{H}_3\text{NHCOPh}$ over Zn dust, whereas the latter yielded a different product. The acridine structure can therefore not be considered for these anhydro oximes. $\text{Ph}_2\text{C NOH}$ also did not give acridine with dehydrating agents. *Anhydrophenyl α -anthraquinonyl ketone oxime* (1.15 g from 2.5 g VIII oxime heated in $\text{AcOH}\cdot\text{H}_2\text{SO}_4$ from 160° to 240° in the course of 10 min.), yellow, m 291° , shows no fluorescence in org. solvents even under a quartz lamp, easily sol. in concd. H_2SO_4 with vivid yellow color, likewise in hot concd. HCl by which it is not changed at 180° , forms a dark orange complex salt when rubbed with anhyd. AlCl_3 , gives a blue vat (sensitive to air) with alk. $\text{Na}_2\text{S}_2\text{O}_4$, yields on distn in H over Zn dust an orange-red sublimatd sol. in org. solvents with blue, in H_2SO_4 with green fluorescence and forming an orange red to brown picrate. The sublimatd is therefore not IX, yellow leaves with green fluorescence, m $197-8^\circ$, which forms an almost black picrate. *Anhydro-*p*-tolyl α -anthraquinonyl ketone oxime*, from the oxime of VII boiled 3 min. in $\text{AcOH}\cdot\text{H}_2\text{SO}_4$, m $236-7^\circ$. *m-Xylol homolog* (III) (yield, 50-60%), yellow, m 240° , insol in NaOH but readily forms a blue-green vat with $\text{NaOH}\cdot\text{Na}_2\text{S}_2\text{O}_4$ at 70° . V (with WILHELM MESSE), from 1 $\text{C}_6\text{H}_4(\text{CO})_2\text{C}_6\text{H}_4\text{COCl}$, *p*-xylene and AlCl_3 in PhNO_2 or CS_2 , light yellow, m 205° , sol in concd. H_2SO_4 with yellow color, changed to green by Al powder, H_2O ppts a blue substance, oxime, m 223° , anhydro oxime (VI) (50% yield), yellow, darkens 255° , m $265-6^\circ$. *Anthraquinone 1-carboxanilide*, light yellow, m $288-9^\circ$, gives anthracene on distn with Zn dust. *m-Xylol- α -aminoanthraquinone* (IV), from $\text{C}_6\text{H}_4(\text{CO})_2\text{C}_6\text{H}_3\text{NH}_2$ and $\text{Me}_2\text{C}_6\text{H}_3\text{COCl}$ in PhNO_2 , m $261-2^\circ$. *2 Methyl anthraquinone 1-carboxanilide*, light yellow, m $287-8^\circ$, forms no anhydro oxime with $\text{AcOH}\cdot\text{H}_2\text{SO}_4$. C. A. R.

Some experiments with fluoranthene. O. KRUBER. *Ber* 64B, 84-5 (1931), cf v Braun and Manz, *C. A.* 25, 1244.—Treatment of 2.4 kg. of a neutral tar-oil fraction (b. about $370-80^\circ$) in an equal amt of purified "solvent benzene" (b. about $140-80^\circ$) at 160° with 300 g. Na gave after 3 hrs stirring a crumbly brown mass which, after thorough

washing with toluene, stirring with ice water, extr with C_6H_6 and distn in *vacuo*, yielded 180 g residue and 800 g distillate, most of which solidified on cooling and on crystn from alc yielded 120 g tetrahydrofluoranthene (I), m 75° , shown by direct comparison to be identical with v. Braun's synthetic product. The part of the original material which had not been attacked by the Na yielded on fractionation 150 g pure pyrene. The I with $Na_2Cr_2O_7$ in AcOH gave an acid, $C_{14}H_{10}O_2$ (probably fluorenone-8-propionic acid), red-yellow, m $137-8^\circ$ (26 g from 5 g I). The I with Na at $140-60^\circ$ readily yields substances which with dry CO_2 give a mixt of CO_2H acids from which was isolated a mono- CO_2H acid, $C_{14}H_{10}O_3$, m 188° , on long heating above its m p or distn with lime the acid regenerates I. The I b₁₁, 363 5° and it is hardly likely that it was already present in the original tar-oil fraction. C. A. R.

Nitrofurfuryl alcohol. HENRY GILMAN AND GEO F WRIGHT *J. Am. Chem. Soc.* 53, 1923-4 (1931), cf. C. A. 24, 5751—Furfuryl acetate and fuming HNO_3 in Ac_2O at -20° give 50-55% of 5(p) NO_2 deriv., m 47° , hydrolysis gives 40-55% of 5(?) nitrofurfuryl alc., b₁ 157° , m 32° , oxidation of the latter with MnO_2 and 50% H_2SO_4 gives nitrofurfural. C. J. WEST.

α -Pyrrolealdehyde. B. EMMERT AND K. DIEHL. *Ber.* 64B, 130-2 (1931)—It was shown recently (C. A. 23, 4698), on the basis of mol wt. data, that α -pyrrolealdehyde (I) exists in soln in an equil. (depending on the solvent) with a form of higher mol

wt. for which was suggested the Jacobson formula $CH \begin{array}{c} \diagup CH \quad C \cdot CH(OH) \quad N \quad CH \\ | \quad | \quad | \\ CH \quad N \quad CH(OH) \quad C \quad CH \end{array} \diagdown CH$ (II).

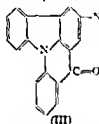
It has now been possible, by benzoylation in petr. ether, to stabilize the 2 forms and sep. them by means of MeOH as their Bz derivs. *N*-Bz deriv. of I, sol in MeOH at room temp., m 90° , mol wt. in Et_2O 171-8, in camphor 198, also obtained exclusively from the Na deriv. of I and $BzCl$ in ligroin, phenylhydrazone, green yellow, m 154° . Dibenzate of II, only slightly sol in cold MeOH, m 178° , mol wt. in camphor 418-25. *N*-*p*-Methylbenzoyl deriv. of I, m $79-80^\circ$, mol wt. in camphor 220, phenylhydrazone, yellow, m 146° . Bis(*p*-methylbenzoate) of II, m $167-8^\circ$, mol wt. in camphor 432. C. A. R.

General method of synthesis for α -substituted pyrrolines and pyrrolidines. LYMAN C. CRAIG, HELEN BULBROOK AND R. M. HIXON *J. Am. Chem. Soc.* 53, 1831-5 (1931), cf. C. A. 24, 1374—A study has been made of the synthesis reported by Cloke (C. A. 23, 2438) for the prepn of α -substituted pyrrolines from $ClCH_2CH_2CH_2CN$ (I). The yields have been increased and the procedure shortened by elimination of $MgClBr$ from the addn. product of the Grignard reagent with I, closing the ring without passing through the intermediate ketimine. α -Ethylpyrroline, b₁ 140° , results in 46% yield, picrate, m 87° , chloroaurate, m 122° . The work of Gabriel and Coleman (C. A. 2, 1425) on the reduction of α -phenylpyrroline (55% yield by the new method) with Sn and HCl has been confirmed, catalytic reduction of this compd. yields an indefinite mixt. γ -Chloropropyl phenyl ketimine acetate is reduced catalytically to 1-phenyl-1-amino-4-chlorobutane, although the latter compd. was not isolated due to the ease with which it splits off HCl to form α -phenylpyrrolidine. α -Benzylpyrroline, b₁₁ $126-8^\circ$, results in 13% yields, picrate, m 89° , chloroaurate, m 125° (decompn.). The yields of these 3 compds. is in the reverse order to the order of reactivity of the Grignard compds. from which they are derived. C. J. WEST.

Behavior of pyrrolidine on dehydrogenation catalysis. N. D. ZELINSKII AND I. K. YUR'EV. *Ber.* 64B, 101-3 (1931)—It was shown recently (C. A. 24, 1110) how easily *N*-methylpyrrole is hydrogenated with Pd under the usual conditions and the resulting *N*-methylpyrrolidine can be dehydrogenated back. It was of interest to det. how pyrrolidine (I) would behave. A com. synthetic pyrrole, b₁ $129-30^\circ$, when purified through the K compd. which was washed several times in H with petroleum ether, dried in H and decompd. with ice, b₁₁ 128.5° , n_D^{20} 1.5093, with H and Pd asbestos at 160° it gave 30% I, b₁ $85-6^\circ$ (in H), n_D^{20} 1.4390, d_4^{20} 0.8618, with Rh as catalyst (which permits of effecting the reduction at 100°), the resulting I b₁ $86-8^\circ$, n_D^{20} 1.3498, I obtained from pyrrolidine (II) with Pd asbestos at 135° and carefully fractionated from BaO b₁₁ $85.5-6.5^\circ$, n_D^{20} 1.4423, d_4^{20} 0.8533. The II, obtained in 35% yield from pyrrole in alc. with Zn-Pd and concd. HCl , b₁₁ $90.0-5^\circ$, n_D^{20} 1.4650, d_4^{20} 0.9017. The I, led at the rate of 4 drops per min. through Pd-asbestos at 300° , yielded pyrrole, b₁ $129-31^\circ$, n_D^{20} 1.4955. C. A. R.

Tervalent nitrogen. II. Carbazofeacridone and its several monosubstitution prod-

ucts. TARO HAYASHI *Bull Inst Phys-Chem Research (Tokyo)* 9, 870-890 (Abstracts 92-3 (in English) published with *Sci Papers Inst Phys-Chem Research (Tokyo)* 15, Nov 278 81 (1930)), cf *C A* 25, 1824.—The following compds were prepd to ascertain the existence of optical isomers of org compds contg a trivalent asym N atom entirely in the ring structure, and having neither an asym C atom nor any other cause for the optical activity. Carbazoleacridone (I), obtained by treating $(C_{11}H_7)NII$ with $ICl_3 \cdot CO_2H$, yellow, m 178.4° (181.9°, cor), whose CO group does not react with $PhNHNH$, showing no ketonic character. 2-Azocarbazoleacridone (II), obtained in 35 g (83%) yield by the addn of 160 cc concd HNO_3 (d 1.38) to 36 g of I in satd. glacial AcOH soln, yellow crystals from $PhNO_2$, m 304.5° . 2-Aminocarbazoleacridone (III), obtained by making alk with NaOH then hot H_2O ext. of the double salt resulting from II, $SnCl_4$ and concd HCl , then treating its helicin compd with alk. KOH (the



three crystals gives poor results), reddish orange crystals from Cl_2 , m 279.40° . HCl salt, orange yellow, m 270.8° . picrate, yellow, m 163.45° . 2-acetyl amino compd, yellow, m $225-6^\circ$. 2-benzoyl amino compd, yellow, m 286.8° . benzoylsulfonamide of I, yellow, m 224.5° . methylamine of I, a secondary amine, reddish orange, m 193.55° . α -bromocamporsulfonate, yellow, m 252.4° . Tartrate, yellow, m 171.2° (decomp), reaction product of III and salicylaldehyde, yellow, m $251-2^\circ$ (decomp). Helicin compd, obtained from III and helicin in hot $EtOH$, an orange ppt decomp at $218.22.5^\circ$, gives, on recrystn from 1 $EtOH$, 2 geometrical isomers, one comparatively insol in 1 $EtOH$, yellowish orange crystals from $EtOH$, decomp $223.5-5^\circ$, $\alpha_D^{20} 0.03^\circ$ (in 0.01% 1 $EtOH$ sol, $l = 1$ dm), another easily sol in $EtOH$, yellow crystals from $EtOH$, m 238.9° , $\alpha_D^{20} 0.02^\circ$ (in 0.100% $EtOH$ sol, $l = 1$ dm). d Hydroxymethylenecampor deriv. of III, obtained in 14.5 g yield by treating 10 g of III in 100 cc glacial AcOH and 200 cc H_2O with 6.5 g d hydroxymethylenecampor in MeOH, yellowish orange crystals from the mixt. of MeOH and H_2O , m $170.5-8^\circ$. Heating 3 hrs with alc potash and addn of alkali or HCl bring about no change, but cold concd H_2SO_4 dissolves it, with pink color, and treatment with H_2O and dil NaOH gives III. There is no indication of the existence of any optical isomer in III. Carbazoleacridonediazonium chloride (V), yellow powder, decomps about 230° . V in 10% cold NaOH, treated with $SnCl_4$ in 50% NaOH and 5% NaOH, gives I. Carbazoleacridone 2-nitro, obtained in 8.0 g yield by treating 10.0 g of V with a hot soln of $CuCN$, yellow crystals from Cl_2 , m $248-9^\circ$, on saponif. it with H_2SO_4 or NaOH, there results carbazoleacridone 2-carboxylic acid, yellow crystals from MeOH, m $288-90^\circ$. Strychnine salt, yellow, decomps $180-4^\circ$. brucine salt, yellow, m $279-70^\circ$. cinchonine salt, yellow, m $220-2^\circ$. The latter 2 salts were optically inactive. III. Configuration of the trivalent nitrogen compound. *Ibid* 10, 1-4 (1931); *Ibid* 15, Nos 283-5, Abstracts (English) 1 (1931).—The structural formula ordinarily given for indazolecarboxylic acid (cf *C A* 25, 1824) selected as an org N compd contg a trivalent asym N atom which is entirely in the ring structure, and an amino or carboxyl radical capable of being combined with active substances, was doubtful and therefore monosubstitution products of carbazoleacridone (cf preceding abstr) were prepd. The expl evidence that the trivalent N atom in 2-aminocarbazoleacridone α methylenecampor and the brucine and guinine salt of carbazoleacridone 2-carboxylic acid does not give rise to the existence of optical isomers and that carbazoleacridone, its 2-nitro and 2-amino derivs, and its 2-carboxylic acid form very weak combinations with MeI and HCl seems to permit of concluding that all 3 bonds of such a trivalent N atom exist in one plane.

K. KONDA

Color of complex diazoles. III. Double quinonoid structure—the real chromophore. GOPAL C CHAKRAVARTI *Proc 15th Indian Sci Cong* 1928, 158, cf *C A* 19, 2493.—All attempts to correlate color with the constitution of compds contg fused pyrrole imidazole or pyridine imidazole ring systems have so far been unsuccessful. Hewitt's rule and its modification by Watson and his co-workers are not sufficiently elastic to explain the color of these nitrogenous bodies. But when the theory of quinonoid structure is extended to heterocyclic compds like pyrrole and pyridine, a ready explanation of the cause of color in these condensed systems is at once available. It is found that all the known colored condensed pyrrole imidazole and pyridine imidazole derivs may be represented as having a double quinonoid pyrrole or pyridine nucleus, whereas the colorless compds either do not possess this structure, or if they do possess it their chromophore effect is neutralized by the presence of other hypsochromic groups. Hence it is suggested that the presence of a double quinonoid pyrrole or pyridine ring

is necessary so that systems contg fused pyrrole- or pyridine-imidazole skeletons may develop visible color. E J C

Action of hydroxylamine upon mustard oils: formation of dianilino-1,2,5-oxadiazole. P. C GUHA AND M N CHAKRADAR *Proc 15th Indian Sci Cong* 1928, 157— NH_2OH reacts with mustard oils to yield unstable intermediate hydroxythiouras, RNHC(S)NHOH , 2 mols. of which are easily decompd even at ordinary temp to yield 1,2,5-oxadiazoles with the sepn of S and water $2 \text{RNHC(S)NHOH} \rightarrow [\text{RNHC(S)NHOH}]_2 \rightarrow \text{RNHC(S)N} \text{O} \text{N} \text{CNHR} + \text{H}_2\text{O} + 2\text{S}$ No details are given E J C

Possible existence of 2-thioimidazole group in insulin. C RUIZ, L SILVA AND L LIBENSON *Compt rend soc biol* 104, 1101 2(1930). *Physiol Abstracts* 15, 530—The 4- or 5-methyl 2-thioimidazole produces hypoglycemia in the rabbit There is no proof that imidazole is the active group in insulin, though histidine has been isolated from it 2-Thioimidazole, the nucleus of ergothioneine, exists in red blood cells It has a hypoglycemic action on the rabbit *per os* 2-Thioimidazole has a similar action Other derms have no such effect. The pos effect is slight and does not suggest that the insulin mol contains the 2-thioimidazole group G G

Dyes derived from oxalyldibenzyl ketone. S A SALETORRE AND GOPAL CHANDRA CHAKRAVARTI *Proc 15th Indian Sci Cong* 1928, 152—Hydroxyimidazoles obtained by condensing phenanthraquinone and acenaphthaquinone with salicylaldehyde in the presence of NH_3 are all colorless crst compds (Japp and Streatfield, *J Chem Soc.* 41, 146(1882) and others), while those derived from substituted salicylaldehydes are colored (Sircar and Sircar, *C A* 17, 2880) It was therefore anticipated that the imidazole from oxalyldibenzyl ketone and salicylaldehyde would likewise be colorless, whereas this compd was obtained as a deep brown crst. product with marked chromophoric properties. This is evidence in support of the observation made recently (C. C. A. 20, 207) that oxalyldibenzyl ketone is a much better chromophore than phenanthraquinone or acenaphthaquinone, although the latter bodies contain condensed benzene nuclei Several other imidazoles derived from oxalyldibenzyl ketone on the one hand and vanillin, nitrosalicylaldehyde, *p*- and *m*- HOCH_2CHO , resorcinol aldehyde, bromosalicylaldehyde, etc., on the other, are also found to be deeply colored No details are given E J C

Some heterocyclic derivatives of biphenyl. J M F LEAPER *J Am Chem Soc* 53, 1891-6(1931)—4- $\text{ClH} \cdot \text{H} \cdot \text{NC}_6\text{H}_4\text{Ph}$ (100 g), slowly added to 700 g of S-Cl_2 at 20° and stirred overnight, then treated with 400 cc C_6H_6 and heated 4 5 hrs at $75-80^\circ$, gives 140 g phenylphenylmethylenethiazolium chloride, brownish red powder, chars about 130° , stirring with H_2O overnight gives the hydrate (I), m 135° (decompn.) I and NaOH give the Na salt of 4-amino-3-mercaptobiphenyl, analyzed as the Zn salt The Na salt and CS_2 in NaOH, refluxed 5 min., give 87% (on the basis of the $\text{PhC}_6\text{H}_4\text{NH}_2$) of 2-mercapto-6-phenylbenzothiazole (II), m 220° Zn salt, Pb salt, yellow, I in NaOH gives bis-6-phenyl 2-benzothiazole disulfide, m 163° II was also obtained from $\text{PhC}_6\text{H}_4\text{NH}_2$, NaSCN and Br in AcOH, which gives 2-amino-6-phenylbenzothiazole, m $227-8^\circ$, heating with concd KOH splits the heterocyclic ring, giving the aminothiophenol which reacts with CS_2 to give II The above Na salt and $\text{ClCH}_2\text{CO}_2\text{H}$ give 94% of 3-amino-5-phenylthioglycolic acid, m $212-25^\circ$ The diazo compd gives a dye with $\beta \text{C}_{12}\text{H}_7\text{OH}$, $\text{C}_6\text{H}_5\text{O}_2\text{N}_2\text{S}$, reddish needles with metallic luster, $\alpha\text{-PhC}_6\text{H}_4\text{OH}$ gives a dye which gives an orange AcOH soln, while that from $\beta \text{C}_{12}\text{H}_7\text{OH}$ gives a bluish red soln. C J WEST

The two isomers of hydroxymethyl phenylcyanopyridine. N TROCCOLI *Ann chim applicata* 21, 41-45(1931)—J Guareschi has shown (*Atti accad Torino* 34) that β -diketones condense with $\text{NCCH}_2\text{CO}_2\text{Et}$ in the presence of NH_4OH to form cyanopyridones. If an asym. diketone is used 2 isomers are obtained. Thus with AcCH_2Cz , the α - as well as α' -phenyl- γ -methyl 3-cyano- α -ketodihydropyridine, m $263-4^\circ$ and 310° resp., are obtained, and are sepd. by their difference in soly in EtOH, the latter being much less sol. Alkali salts form readily by treating the EtOH suspension with 10% alkali. The salts of the heavy metals can then be prepd. from the alkali salts by adding sol. metal salts, by metathesis. In this way the K, Na, Ba, Cu, Ni and Co salts of each isomer were prepd. A W CONTIERI

The condensation of pyridine- and quinolinecarboxylic acids with ammo acids. HANS MEYER AND ROBERT GRAY *Biochem. Z.* 229, 154-63(1930)— $\text{CH}_3(\text{NH}_2)\text{CO}_2\text{Et}$ and 2 $\text{C}_6\text{H}_4\text{NCOCl}$ give Et 2-pyridylamiracetate (I), m 71° . Sapon. gives the free acid, m $165-6^\circ$ (cf Sendju, *C A* 22, 602), itsamide, m $184-5^\circ$, is obtained from I and NH_4OH The following compds. were prepd. similarly. 3-pyridylamiracetic acid, m $237-8^\circ$ (decompn.) (amide, m $193-5^\circ$, Et and Me esters, m $54-5^\circ$ and $67-8^\circ$.

resp.), 4-pyridylaminoacetic acid, m 221-2° (decompn) (Et ester, m 89-90°; amide, m 227-8° (decompn)), 2-quinolylaminoacetic acid, m 183-4° (decompn.) (Me ester, m 100°; amide, m 231-2° (decompn)), 4-quinolylaminoacetic acid, m. 244-6° (decompn) (monohydrate, decompn 110°, amide, m 226-7°; Et ester monohydrate, sinters 110°, m clear 120°), 2-phenyl-4-quinolylaminoacetic acid, m. 216° (Et ester, m 140-1°, amide, m 213°), 3-pyridylglycylglycine, m 232° (decompn) (Et ester, m 144-5°), N (3-pyridyl)anthranilic acid, m 258-9° (Me ester, m 127°), N (2-pyridyl)anthranilic acid, m 171-2° (Me ester, m 84-5°), N-(3-pyridyl)-anthranoylanthranilic acid, m 222° (Me ester, m 181° (decompn)) A. F. SHEPARD

Anhydrides of pyridine- and quinoline-carboxylic acids. RODERICH GRAY. *Biochem Z* 229, 164-8 (1930), cf C A. 23, 837.—Nicotiny chloride and Na nicotinate refluxed 1 hr in $\text{C}_2\text{H}_5\text{I}$ give nicotinic anhydride (I) m 122-4°. Isonicotinic anhydride (II), prep'd similarly, m 103-4° and carbonizes 160°. I and II are only slowly hydrolyzed. Na cinchoninate and cinchoninyl chloride heated 0.5 hr at 170° give cinchoninic anhydride b, about 240-50°, m 215° (decompn). Anhydrides could not be obtained from quinaldic acid or 2-pyridinecarboxylic acid. A. F. SHEPARD

α -Piperidinobenzylacetophenone, 1 P KOHLER and W. F. BAUER. *J Am Chem Soc* 53, 1934-8 (1931).—K. and ADDINALL (C. A. 24, 5030) expressed the opinion that a peculiar red piperidine deriv (Watson, *J. Chem Soc* 85, 1322 (1904)) is an unsat'd. ethylene oxide, $\text{PhCH}_2\text{C}(\text{C}(\text{N}(\text{C}_2\text{H}_5)_2)\text{O})_2$ (I); however, it is an unsat'd

ketone, $\text{PhCH}_2\text{C}(\text{C}_2\text{H}_5)_2\text{N}(\text{C}_2\text{H}_5)_2$ (II), as by DUFRASSE MOURTU (C. A. 21, 3051). Hydrogenation gives α -piperidinobenzylacetophenone (III), pale yellow, m 81°, also obtained from $\text{PhCH}_2\text{CHBrBr}$ and piperidine, the oxime m 141-2°, it was not possible to regenerate the ketone from the oxime. III and PhMgBr give α, γ -triphenyl- β -piperidinopropanol, m 148-9°, pyrolysis of the carbinol gives $\text{PhCH}_2\text{CH}_2\text{N}(\text{C}_2\text{H}_5)_2$ and Ph_2CO . I and PhMgBr give α -piperidino- β, δ -diphenylpropiophenone, yellow, m 163°; further action of PhMgBr gives α -diphenyl- β -piperidino- γ, γ -diphenylpropanol, m 136-7°, the chloride m 162-4°. C. J. WEST

The bromination of methylarsepidine. The existence of a radical with quadrivalent arsenic. RENÉQUE V. ZAPPI AND HELVÉCIO DEGIORGI. *Bull soc chim* [4], 49, 366-71

(1931), cf C A 10, 2874.—The bromination of methylarsepidine, $\text{CH}_3(\text{CH}_2)_3\text{CH}_2\text{AsMe}$ (I) is exothermic and forms the dibromide, $\text{CH}_3(\text{CH}_2)_3\text{CH}_2\text{AsBr}_2\text{Me}$ (II). On

adding about 50 cc. of 2 N HBr in CCl_4 (contg 0.1 atom Br) to a cold soln. of 16.1 g of I (0.1 mol), the reaction ceases with the production of dibromomethylarsepidine,

$[\text{CH}_3(\text{CH}_2)_3\text{CH}_2\text{AsBrMe}]_2$ (III). By allowing this soln to attain room temp. and on adding 50 cc. of 2 N HBr, II was ppt'd. The white cryst. material (III), m. 60° (decompn.), is very hygroscopic, sol. in warm CCl_4 to a yellow soln. which is decolorized

on exposure to the air, pptg the peroxide, $[\text{CH}_3(\text{CH}_2)_3\text{CH}_2\text{As}(\text{MeBr})\text{O}]_2$. It would seem that III is capable of dissociating into the free radical bromomethylarsepidyl

$\text{CH}_3(\text{CH}_2)_3\text{CH}_2\text{As}(\text{MeBr})$, contg As in the quadrivalent state. Though there are no previous reports on quadrivalent As, Blöcke and Smith (C. A. 23, 3908) have reported bivalent As in the tetraaryldiarsyls. When II is heated to 60-80° at 6-8 mm, or when a soln. of II in CCl_4 is distilled to a small vol., MeBr is evolved, with formation

of bromoarsepidine, $\text{CH}_3(\text{CH}_2)_3\text{CH}_2\text{AsBr}$, a red oil which is converted on bromi-

nation into arsepidine tribromide, $\text{CH}_3(\text{CH}_2)_3\text{CH}_2\text{AsBr}_3$ (IV), m. 102°. IV is de-

comp'd by H_2O into HBr and $\text{CH}_3(\text{CH}_2)_3\text{CH}_2\text{As}(\text{OH})\text{O}$, and on heating gave gaseous

hydrocarbons and AsBr, but not $\text{CH}_3\text{CH}_2\text{CHCHCH}_2\text{As}$ as was hoped. Attempts

to form the pentabromide, $\text{CH}_3(\text{CH}_2)_3\text{CH}_2\text{AsBr}_5$, were unsuccessful. C. R. A.

Pyrimidines. CXXII. Improved methods for the synthesis of orotic acid. THOMAS B. JOHNSON and ELMER F. SCHROEDER. *J Am Chem Soc* 53, 1983-94 (1931); cf C. A. 25, 1830.—Uracil-4-carboxylic acid (orotic acid) (I), known to occur in milk, is most easily prep'd in quantity by oxidation of 2-thiouracil-4-aldehyde (II) or uracil-4-aldehyde (III) with CrO_2 . Details are given for the prep'n. of $(\text{EtO})_2\text{CHCO}_2\text{Et}$, $(\text{EtO})_2\text{C}$

CH₃COCH₂CO₂H, the di-*i*t acetal (IV) of II, II, and the di-*i*t acetal (V) of 2 ethylmercaptouracil-4 aldehyde (VI). Oxidation of II gives 76% of I, identical with the natural product; VI gives 55% of I; IV gives 68% of I, V gives 53% of I. IV, oxidized with 3% H₂O₂ in 20% NaOH, gives a nearly quantitative yield of the di-*i*t acetal of III; oxidation of the latter gives a good yield of I. A quantitative yield of III is obtained by the careful hydrolysis of corresponding di-*i*t acetal; III contains 1 mol H₂O, lost on heating 1 hr at 120°. It is easily oxidized to I.

Cyclic ammonio ketones and acid chlorides of the quinoxaline series. R. A. OGGS, JR., and T. W. BERGSTRÖM. *J. Am. Chem. Soc.* 53, 1816-13 (1931), cf. C. A. 25, 957—2,3-Dichloroquinoxaline (I) is a deriv. of an ammonio oxalyl chloride, the method of prepn., from 1-Cl₂ and 2,3-dihydroxyquinoxaline (previously shown to be an ester of a mixed aquoammonio oxalic acid), is in accord with this view. I also reacts with (C₆H₅)₂NH₂ to give *o*-phenylenedihydroquinoxaline, leaves, m above 750° ale solns show a faint fluorescence. I also reacts with 2 equivs of MeMgI to give 64% of 2,3-dimethylquinoxaline (II) and with PrMgI to give 30% of 2,3-dipropylquinoxaline (III) m. 42.9° (cor). I did not react with EtMgBr even after prolonged heating. II is a deriv. of an ammonio diacetyl—this is seen in its synthesis from I and MeMgI and from *o*-C₆H₄(NH₂)₂ and di-Ac and in its behavior on reduction. Further, Et₂NH gives a dibenzal deriv. II does not react with Grignard reagents or with HCN. With 2 equivs of KNH₂, II gives a di-K salt, greenish yellow, which yields with 1 *l* III (70% yield). 2,3-Diphenylquinoxaline (IV) is a deriv. of an ammonio benzil as shown by its synthesis and behavior on reduction and oxidation. IV does not react with Grignard reagents or with HCN, with KNH₂, it appears to undergo a benzilic acid rearrangement giving 30% of what is assumed to be 1,2-dihydro-2,2-diphenyl-3-aminoquinoxaline, m 287° (cor). It does not react with HNO₃ or with I (60%) HCl at 150°. These reactions constitute 2 new methods for the synthesis of 2,3-dialkylquinoxalines. C. J. W.

Synthesis of substituted thiazines. S. KRISHNA and MITTER S. JAIN. *Proc. 15th Indian Sci. Cong.* 1928, 153-4.—The methods of prepn. of substituted thiazines hitherto described give only the symmetrically disubstituted derivs. and are applicable only in a very limited no. of cases. An attempt has therefore been made to synthesize mono-, di- or polysubstituted thiazines by a method that should be applicable to C₆H₅ derivs. containing acid or basic substituents. Thiazine was synthesized by a method utilizing the reactivity of the halogen atom in negatively substituted chlorobenzene. 6,3-Cl(O₂N)₂C₆H₃SO₂H condenses with aniline to form 4,2-O₂N(HO₂S)₂C₆H₃NHPh. This is readily sol. in H₂SO₄. If the blue soln. obtained is immediately diluted with water, 3-nitrophenothiazine S oxide is pptd., but if the soln. is kept for half an hr. before diln., 3-nitrophenothiazine (I) is obtained and SO₂ evolves. I is readily reduced to 3-amino phenothiazine from which, through the diazo compd., phenothiazine is obtained. A blue color is produced when a soln. of I in AcOH is treated with HCl gas. The production of this blue color has been regarded as evidence of the existence of quinonoid dithionium salts. Attempts have been made to isolate such salts but these have not been successful, probably on account of the influence of a nitro group. E. J. C.

The hydrogenation of methylene blue in the absence of enzymes. PRITZ LARSEN and ERICH MOISNER. *Biochem. Z.* 232, 209-17 (1931).—Glycocholic as well as albumin inhibit the effect of dil. NaOH in deepening the color of methylene blue and lead thus to the formation of stable light blue solns. Since this effect can also be produced by means of isorg. buffers, it is concluded that the ammoniacal buffers probably produce this effect through their buffer action. The hydrogenation of methylene blue by urea or guanidine derivs. manifests certain regularities. The addn. of CO groups to the —NHCONH₂ or —NH(C=NH)NH₂— promotes the reduction, whereas CH₃ groups inhibit it. Urea and guanidine themselves are inactive, nor do the groups —CONH₂— or —CO₂— occasion reduction. Opening the ring diminishes or even abolishes the reducing action. The reducing effect diminishes along the series: alloxan > parabanic acid > allantoin > hydantoin also hydantoin > isyllantoin acid, creatinine > creatine. This shows the stimulating effect of CO groups or of the cyclic structure. S. M.

Isomerism. Ring closure of *o*-thiocarbamidobenzole acids. TRJENDRA NATH GHOSH. *J. Indian Chem. Soc.* 7, 941-4 (1930).—When anthranilic acid and PhNCS are allowed to react in the cold, *o*-phenylthiocarbamidobenzoic acid, *o*-RNHCSNH(C₆H₅)CO₂H (I), (R = H), and 4-keto-2-thio-3-phenyl-1,2,3,4-tetrahydroquinazoline, C₆H₅NHCSNHCO₂H (II), are obtained. Anthranilic acid has now been condensed

with various mustard oils in hot alc. and in each case only the corresponding quinoxaline has been obtained, showing that the arylthiocarbamidobenzoic acid very easily passes into the quinoxaline with loss of water. When II is heated with concd. H₂SO₄

for 3-4 hrs. at 125-30° it is isomerized to a compd. of the type $C_6H_4 N \cdot C(NH_2) S CO$.

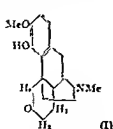
2-amino 3-keto-4,5-benzo 1,3-thiazine (III), insol in alkali but sol in acids. III is also obtained from I by the same treatment. The allyl analog gave, however, a 3-free compd with H_2SO_4 but a compd similar to III when HCl was used. II in above soln, m 254-2°, disulfide (by addn of I_2 in KI to II in $AcOH$ and gentle warming followed by diln), m 254-2°, *p*-tolyl analog, m 310°, *o*-tolyl analog, m 254-70°, disulfide, m 215°, 1,3,4-xylyl analog, m 259-60°, allyl analog, m 206-7°. III, m 184-5°, *p*-tolyl analog, m 235°, allyl analog, m 115°. R. C. FIDLERFIELD

Thiodiazines. VII. Condensation of ethyl chloroacetate with thiosemicarbazides. PRAFULLA KUMAR BOSE AND HIRENDRA KUMAR NANDI. *J. Indian Chem. Soc.* 7, 961-4 (1930), cf *C. A.* 25, 1532.—Thiosemicarbazide and its 4 substituted derivs condense with $ClCH_2CO_2Et$ to give in good yields thiodiazines, $RNH \cdot C \equiv N \cdot N \cdot C(OH)CH_2S(H)$, sol in warm aq alkalis and acids. The indifference of

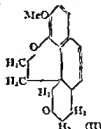
these compds to aldehydes and ketones indicates a thiodiazine rather than a thiazole formula. Mol proportions of the thiosemicarbazide and $ClCH_2CO_2Et$ were refluxed in $EtOH$ for 1 hr. The cryst condensation product sepd during the reaction. 2-Amino 5-hydroxy 1,2,4-thiodiazine ($R = H$ in I), m 254° (decompn) (Ac deriv, m 205°, Bz deriv, m 260°, addn compd with phenylthiocarbamide, m 195°), 2-Me homolog, m 282° (decompn) (Ac deriv, m 194°), 2-Et homolog, m 225° (Ac deriv, m 165°), 2-n-Bu homolog, m 210° (Ac deriv, m 199°, mono-Me deriv, by refluxing with slightly more than 1 mol MeI and 1% KOH in MeOH, m 135°), 2-phenylamino analog, m 184° (Ac deriv, m 172°, mono-Me deriv, m 265°), 2-*p*-tolylamino analog, m 185° (Ac deriv, m 210°), 2-*p*-tolylamino analog, m 195° (Ac deriv, m 218°). R. C. FIDLERFIELD

Calycanthine. II. The degradation of calycanthine to *N*-methyltryptamine. RICHARD H. F. MANXEP. *Can. J. Research* 4, 275-82 (1931).—Degradation of calycanthine by benzoylation with subsequent treatment with alkali gave a well-defined cryst. substance, benzoyl *N*-methyltryptamine. The identity of this substance was established by both analysis (C 77.14, 76.87, H 6.58, 6.40, N 9.90, 9.95, mol wt 293, 305) and synthesis. *m*-Chlorobenzoyl and *p*-nitrobenzoyl *N*-methyltryptamine were prepd., the former consists of short stout prisms, m 153°, the latter of golden yellow rectangular plates, m 131°. Benzoyl *N*-methyltryptamine formed synthetically from BzCl and *N*-methyltryptamine was recovered as colorless acicular needles, m 117°. J. W. SMILEY

Sinomenine. XXIV. Degradation of sinomenine to *l*-thebenone. KAKUJI GOTO WITH REIKICHI INABA AND HIRORO SHIRAI. *Ann.* 485, 247-57 (1931), cf *C. A.* 25, 2432 and following abstr.—Demethoxydihydrosinomenine MeI, heated with 25% KOH 10 min, gives 80% of des *N*-methyl-desmethoxydihydrosinomenine, $C_{17}H_{19}O_2N$ (I), m 182°, $[\alpha]_D^{25} = -54.94^\circ$ ($CHCl_3$), concd H_2SO_4 gives a deep red soln, while this has the same m p as the *N*-Me base from dihydrothebanone, the rotation is different. 1 MeI gives with KOH 44% of dehydrothebenone, $C_{17}H_{17}O_2$ (II), m 104°, $[\alpha]_D^{25} = -206.87^\circ$ ($CHCl_3$). Catalytic reduction of I gives nearly quant. the dihydro deriv, m 156.5°, $[\alpha]_D^{25} = 67.82^\circ$, methiodide, m 226-9°, with KOH the latter yields *l*-thebenone (III), $C_{17}H_{15}O$, m 134°, $[\alpha]_D^{25} = -78.60^\circ$, oxime, m 204.5°; III also results by the reduction of II.



(I)

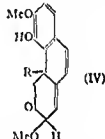
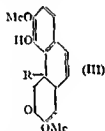
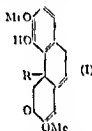


(II)

C. J. WEST

Sinomenine and disinomenine. XXV. Three different sinomeninemethines. KAKUJI GOTO AND HIRORO SHIRAI. *Bull. Chem. Soc. Japan* 6, 79-87 (1931), cf

preceding abstr—Sinomenine MeI and 2 mols 2% NaOH, boiled 1 min, give *sinomenine achro-methine* (I, $R = Me_2NC_6H_4CH_2-$) (previously termed *N*-methylulhydrasinomenine, *C. A.* 21, 1166) the name arises from the very faintly yellow halochromy in concd H_2SO_4 and to avoid confusion with the methylmorphinethines), 1st purified through the Na salt, m 170°, $[\alpha]_D^{25}$ 72.58° ($CHCl_3$), easily sol in H_2O , III salt, m 115–8°, the HCl and HBr salts are very hygroscopic. *methiodide* (II), m 212°, $[\alpha]_D^{25}$ –37.00° (H_2O) crizme, decomp 204.5°, decompn of I with 66% KOH or Bz_2O gives the same sinomenol or Bz deriv as with sinomenine itself. I on standing for years gives 10% of *sinomenine roseo methine* (III) (previously termed β -sinomenine methine), also obtained by decompn of sinomenine-MeI with 5 mols 5% NaOH, yellow, m 163°, gives an intense red halochromy in concd H_2SO_4 and a yellow red in concd HCl $[\alpha]_D^{25}$ 135.70° ($CHCl_3$), difficultly sol in H_2O , *methiodide*, m 277° (decompn), $[\alpha]_D^{25}$ 18.23° (H_2O) I with 10% NaOH gives 20% of *sinomenine violeo methine* (IV) (previously termed the α deriv), m 172.3°, almost insol in H_2O , $[\alpha]_D^{25}$ 471.75° the *methiodide*, m 209°, $[\alpha]_D^{25}$ 37.336° (H_2O), results in almost quant yield by the action of cold 10% NaOH upon II; both compds give an intense blue halochromy in concd H_2SO_4 or fuming HCl, heating the MeI deriv with 3.3% NaOH for 1 hr gives sinomenol. The proposed structures for these 3 compds are discussed



C. J. WIST

The structure of yohimbine. P. MENDELIC *Pharm. Weekblad* 68, 257–71 (1911) — The alkaloid yohimbine $C_{17}H_{19}O_2N$, which is the Me ester of yohimbic acid and contains a OH, was mixed with Se and heated in a bath at 300°. Dehydrogenation occurred with evolution of H_2Se . 1 xtn of the reaction mass with CaH_2 yielded a cryst base yohyryne (I), $C_{17}H_{17}N$, m 217°, the salts of which have a violet blue fluorescence. From the CaH_2 mother liquor a more sol base *dihydroyohyryne* (II), $C_{17}H_{19}N$, m 170°, was obtained. The residue insol in CaH_2 was extl with IOH and yielded *kelyoyhyryne* (III), $C_{17}H_{17}ON$, m 324°. Fusion of III with KOH at 350° gave a cryst base $C_{17}H_{19}O_2N$, m 258°, and an acid, m 144°, which was identified as *hemimellitic acid*. Yohimbine probably contains a cyclohexene ring. The first thermal reaction product of the acid is apoyohimbic acid in which the OH and an adjacent H have split out as H_2O . Dehydrogenation by Se then splits off CO_2 from the CO_2H and removes 4 H, leaving I and the intermediate II. The products are the same whether yohimbic acid or its Me ester (yohimbine) is used. In the formation of III the dehydrogenation of apoyohimbine is attended by loss of H_2O instead of loss of CO_2 . The base $C_{17}H_{17}N$ obtained by dry distn of yohimbine differs from I by 6 C and 6 H. It is possible that a hydrogenated *o*-xylene ring is present which is broken up in the distn leaving 2 H at the places of attachment, while the same ring is rendered stable by the Se dehydrogenation. Dimethylindole, another decompn product of yohimbine, may possibly represent the structural portion from which the hemimellitic acid was obtained by oxidation.

A. W. DOX

Diastereoisomerism. VI. Configuration of the morphine alkaloids. HERMANN FIMOR *Helv. Chim. Acta* 13, 1035–56 (1930), cf *C. A.* 24, 2115 — It is deduced from observations recorded in the literature and from the detn of the rotation of derivatives that the rotation of the 5 asym C atoms in morphine is $C_5(-)$, $C_6(-)$, $C_7(-)$, $C_8(+)$, $C_9(+)$, numbered according to Gulland and Robinson (*C. A.* 17, 2594). Based on these conclusions regarding morphine, it was correctly predicted that thebaine would be *l*- and dihydrothebaine *d* rotatory. The isomers of morphine and of codeine prove to be exceptions to this, e. g., α -isomorphine rotates more strongly *levor* than morphine although it would be expected to rotate less strongly *levor* if these compds were true epimerides. The explanation that all 5 rings may not be in the same plane

may answer here. If a model for morphine is constructed using planar rings, the quaternary asym C atom (C_{11}) prevents all 5 rings from being in the same plane, ring III or V is inclined against the plane of the other 4 rings. In apomorphine or morphothebaine, when C_{11} is not asym, all 5 rings may be arranged in the same plane. $[M]_D$ (in water unless stated otherwise) are recorded as follows: morphine $\cdot HCl$ ($3H_2O$) -370° , codeine $\cdot HCl$ ($2H_2O$) -402° , thebaine HCl (H_2O) -574° , apomorphine HCl -149° , morphothebaine HCl -146° , codeine MeI -362° , α -6-acetylmorphine HCl -593° , diacetylmorphine HCl -619° , dipropionylmorphine, m. 107° , -685° , in alc., dipropionylmorphine -671° to dil. HCl (HCl salt, m. 210° (not sharp), -613°), dihydrocodeine -459° (II tartrate -588°), dihydrothebaine -623° in dil. HCl .

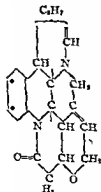
A. WIENEN

Cinchona alkaloids. VII. Sulfonation of quinone and hydroquinone. G. GIMESA AND M. OESTERLIN. *Ber* 64B, 57-61 (1931); cf. *C. A.* 19, 1425, KITASO and GOTO *C. A.* 25, 1532.—The striking ease with which the hydroquinone-sulfonic acids are formed and saponified led to the suspicion that sulfonation does not occur in the nucleus and that the supposed SO_3H acids result from esterification of the alc. group and are really acid sulfates. Sulfonation occurs only when the alc. group is present, if it is acylated, either the acyl group is saponified and the sulfonic acid is formed or, under the proper conditions, neither sapon nor sulfonation occurs. Accordingly, hydroquinone chloride and deoxyhydroquinone form no sulfonic acid, while acetyl and benzoylhydroquinone undergo a quant. exchange of acyl radicals (*Umsetzung*). On the other hand, it is not possible to obtain from hydroquinonesulfonic acids the corresponding acylsulfonic acids, even with such reactive agents as $AcCl$ or $ClCOEt$. It had been shown that treatment of freshly diazotized 5'-aminohydroquinone with freshly prep'd. Cu paste reduces the alc. group simultaneously with the elimination of N , giving methylhydrocuprean. Diazotized 5'-aminohydroquinonesulfonic acid behaves in the same way, splitting off the sulfonic group (at room temp.) and likewise giving the cuprean. Because of the possibility of rearrangement into the iso base, quinone with conc'd. H_2SO_4 gives no homogeneous product. The rearrangement can be completely prevented, however, by using Ac_2O , the resulting *quinonesulfonic acid* (I) is difficultly sol. in H_2O , can be converted back into quinone, adds 1 mol. Br_2 to form *quinonesulfonic acid dibromide* (II) (which can be saponified to quinone dibromide) and is oxidized by $KMnO_4$ to *quinonesulfonic acid* (III), which forms gussetine with HCl . The quant. formation of I in Ac_2O from quinone bisulfate would suggest an intermediate formation of acetylquinone, which then reacts with the H_2SO_4 , in expts. with acetylquinone under analogous conditions, however, almost 90% was recovered unchanged. With conc'd. H_2SO_4 quinone in the cold gives chiefly *hydroxyhydroquinonesulfonic acid* (IV) (mixture of the α - and β isomers), along with isoquinone- and mequinonesulfonic acids. I, from quinone bisulfate in Ac_2O at room temp. or, more simply, from neutral quinone sulfate in Ac_2O with the amt. of H_2SO_4 necessary for the formation of the bisulfate, darkens around 230° , decomps. 237° , gives a positive thalleoquinone reaction, rapidly decolorizes $KMnO_4$, $[\alpha]_D^{25}$ 88.86° ($NaOH$), *mono-HCl* salt, needles with 5 H_2O , m. 205° , is quantitatively converted back into quinone by boiling 25% HCl . II, m. 232° . III, very hygroscopic, m. $221-5^\circ$. IV has not as yet been crystd.; the H_2SO_4 soln. shows blue fluorescence and only slowly decolorizes $KMnO_4$; the thalleoquinone reaction is positive; it forms no NO deriv. but yields with $Ac-O$ a di-Ac deriv. C. A. R.

The molecular structure of strychnine and brucine. ROBERT ROBINSON. *Proc. Roy. Soc. (London)* A130, 431-52 (1931).—Bakerian Lecture, May 1930. A summary of the chemistry of strychnine and brucine (dimethoxystrychnine). Evidence is given to prove the occurrence of 5 groups of atoms in strychnine. These 5 groups are combined, giving a structure containing 5 fused rings. The disposition of the atoms of the group C_8H_7 cannot be predicted but they must be added so as to introduce 2 new rings. The structure of strychnine is represented by the accompanying formula, brucine contains MeO groups instead of H at the 2 positions marked with an asterisk.

JANET E. AUSTIN

Synthesis and pharmacodynamic action of a homoisopapaverine (and landanosine). I. Synthesis of β -2,3-dimethoxyquinoline. FRIGYES (FRITZ) KONIG. *Malacanth. és Természettud.* *Értésl.* 47, 779-86 (in German 787) (1930).—Veratrole in glacial $AcOH$ is nitrated with 1.5 HNO_3 , the resulting nitroveratrole is reduced with Sn to aminoveratrole, then

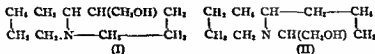


Értésl. 47, 779-86 (in German 787) (1930).—Veratrole in glacial $AcOH$ is nitrated with 1.5 HNO_3 , the resulting nitroveratrole is reduced with Sn to aminoveratrole, then

treated with fuming HCl and paraldehyde, 0.25 vol alc and 0.25 vol ether is added and the mixt cooled to -10 – 15° . The HCl salt of the base which ppts is water-sol. and has an anesthetic action when put on the tongue. S S DE FINÁLY

Problems of the technical manufacture of papaverine (and landanosine). FRIGYES (FRITZ) KONEK. *Matematik és Természettud Értéslát* 47, 788–93 (in German 794) (1930). —For the time being neither the method of Pictet, that of Rosenmund nor that of Pictet-Spáth seems to be economical for papaverine manuf. S S DE FINÁLY

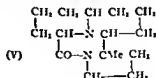
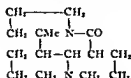
Constitution of lupinine. I. K. WINTERFELD and F. W. HOLSCHNEIDER. *Ber.* 64B, 137–50 (1931). —In the course of their investigations on the lupine alkaloids (C. A. 24, 4787), W. and H. have confirmed one of Karrer's suggested formulas for lupinine (I) but have also found that I is not homogeneous, being accompanied by a structural isomer (II), for which the name *allo-lupinine* is suggested. These conclusions are based on the following series of reactions: $C_{15}H_{21}ON$ (I) ($-H_2O$) \rightarrow anhydrolupinine (III) ($+ H_2$, Pd) \rightarrow lupinane, $C_{15}H_{23}N$ (IV) ($+ BrCN$) \rightarrow bromolupinane, $C_{15}H_{23}N$ (IV) ($+ H_2$, Pd) \rightarrow lupinane, $C_{15}H_{23}N$ (IV) ($+ HCl$) \rightarrow $C_{15}H_{23}N$ (VII) ($-6 H_2$) \rightarrow $C_{15}H_{23}N$ (VIII). In spite of the most careful purification, these intermediate products did not seem to be homogeneous, as indicated by the not quite sharp m or decompn ps of individual salts and of N-derivs. of VII. VII has a secondary cyclic N atom and all the properties of a piperidine deriv., VIII a tertiary cyclic N atom and the properties of a C_8H_7N homolog. Oxidation of the VIII in H_2O with varying amts of $KMnO_4$ gave α, α' - (IX) and β, α' - $C_8H_7NMeCO_2H$ (X), α, δ - C_8H_7N - $(CO_2H)_2$ (XI) and a mono- CO_2H acid $C_8H_7NO_3$ (XII) giving the Fe salt reaction characteristic of α -pyridinecarboxylic acids. The extraordinarily slow oxidation of the side chain C_8H_7 in XII and the failure of I and IV to give a pyrrole reaction indicate that this C_8H_7 group contains a straight C chain and that XII is α -butylpyridine- α' -carboxylic acid. It follows that VIII is a mixt of α -butyl- δ - and δ' -methylpyridines, and VII a mixt. of the corresponding piperidines, that I and II have the structures shown in the accompanying formulas and IV is a mixt. of the corresponding compds. with Me instead of CH_2OH . Yields of the oxidation products of VIII indicate that in the product isolated by W. and H. from the seeds of the yellow lupine I predominates. Since no trace of α -picolinic acid could be detected in the oxidation products, cleavage by the action of $BrCN$ on IV must take place almost exclusively in the non-substituted ring. In the light of the structure thus established for I, there can hardly be any doubt that the quant. rearrangement of the l rotatory I by Na in boiling C_6H_6 into the d -rotatory isolupinine (Krieg, *Diss. Marburg*, 1928) is a case of *cis-trans* isomerism, I being the energy richer *cis* form. V (yield, 90%), yellow to light red oil, is very slightly sol. in 15% HCl. In the prepn of VI is formed as by product a base $C_{15}H_{23}N$ (XIII), isolated as the chloroplatinate, orange, m. $217-8^{\circ}$ (decompn). The VI (yield, 85%) is a light yellow mobile oil of garbolic odor. VII, b_p $101-2^{\circ}$ (decompn), b. $169-71^{\circ}$ (decompn), gives the characteristic piperidine reaction (blue-violet color) with the Levin reagent, HCl salt, m. $151-3^{\circ}$, easily sol. in H_2O with weak acid reaction, is optically inactive (0.0508 g. in 10 cc 96% alc. in a 10-cm tube), HI salt, readily becomes discolored in the light, HBr salt, m. $168-70^{\circ}$, N-Bz and N-*p*-nitrobenzoyl derivs., yellow oils practically insol. in dil acids and H_2O (the Bz deriv. requires 4–5 hrs. heating with concd HCl at 150° for quant. hydrolysis), Ac deriv., liquid, β -naphthalenesulfonyl derivative, leaflets from aq. MeOH, m. $86-7^{\circ}$, granular and of higher



m. p. from petroleum ether, of lower m. p. ($57-9^{\circ}$) from Me_2CO . The HCl salt heated with concd aq. KOH and then treated with $BzCl$ gives a Bz deriv., $C_{15}H_{23}NCONBz$, m. $169-70^{\circ}$, of the half-sapon. VI, also obtained by moderate sapon. of VI with alc. HCl at 85° . XIII is also obtained from VI with H_2 in 96% alc. with Pd $CaCO_3$. VIII, b_p $110-2^{\circ}$, b. $153-4^{\circ}$ (decompn), chloroplatinate, orange-to yellow red, m. $190-1^{\circ}$, decompn. $192-4^{\circ}$; chlorourate, m. $46-7^{\circ}$. XII was isolated as the dark violet Cu salt; chloroplatinate, yellow, m. $154-5^{\circ}$. Cu salt of IX, light blue crystals with 1 H_2O , m. $253-5^{\circ}$ (decompn); HCl salt, m. $202.5-3.5^{\circ}$ (decompn); free acid, m. $95-6^{\circ}$, identical with a product synthesized from 2,6- $C_8H_7NMe_2$. X was isolated as the chloroplatinate, orange-red crystals with 2 H_2O , m. $159-90^{\circ}$ (decompn). Cu salt of XI, dark blue prisms with 2 H_2O or prismatic needles with 1 H_2O , darkens and evolves gas $190-5^{\circ}$, m. 229° (decompn), identical with synthetic quinolinic acid; synthetic α, α' -

pyridinedicarboxylic acid, from 2 G $C_6H_4NMe_2$, m $235-6^\circ$ (decompn) and depresses the m p of XI $10-2^\circ$ C A R.

Lupanine. K WINTERFELD AND A KNEUER *Ber* 64B, 1938 (1931) —Lupanine (I), the main alkaloid of the blue lupine, differs from sparteine (II), an alkaloid of the yellow lupine, only by having an O instead of 2 H atoms. As this O atom is present neither as HO nor as an ether, it was concluded that it is present as a C O group, and since I, although it is bitertiary, is a mono-acid base, the C O group must form a lactam ring with a N atom, a view confirmed by Clemo and Leitch who with lumine III and red I' were able to replace the O in I with H and obtain in good yield a desoxylupanine isomeric with II. W and K by the same reaction under somewhat different conditions (heating 50 instead of 36 hrs and raising the temp to 210°) have been able to break down the 1 mol and obtain 2 cleavage products $C_{11}H_{17}N$ (III), yellow oil of a hay like odor, b_p $85-6^\circ$, identified through its chloroaurate, m $143-4^\circ$, and picrate, m $164-5^\circ$, with β lupanine, and a product (IV) b_p $147-7^\circ$, which yielded no cryst salts but distinctly gave the pine splinter pyrrole reaction. Ten of the C atoms in I must therefore have the same structure as in β lupanine and the other five are probably in the form of a pyrrolidine nucleus. Oxidative degradation of II gives a methyl pyrrolidine, assuming a similar ring system in I, the (still purely hypothetical) structure V may be assigned to I. Direct comparison of III with a sample of the decompn product of matrine (Kondo, C A 22, 1212) proved that the 2 compds are identical. Kneuer (*Diss Freiburg Br* 1929) obtained from the mother liquors of lupanine a new alkaloid, *lupandine*, isomeric with I, which, in view of the facts presented in the preceding abstr, is probably the structural isomer VI. Contrary to Thoms and Bergerhoff, W and K were able to effect a quant cleavage of I with $BrCN$ by using boiling CaH_2 and excluding moisture. The well-crystd bromocyanamide (VII) after removal of the Br and sapon of the CN group gave an oily base (VIII) whose secondary cyclic N atom was readily benzoylated but the ring in the Bz deriv could not be broken open with PBr_3 or PCl_5 , at the high temp required, considerable halogen substitution took place, exhaustive methylation also resulted in only 15% cleavage, boiling the quaternary NH_4 base chiefly split off MeOH. Chloroplatinate of III, red, decomps 217° . IV, b_p $72-3^\circ$, mol wt in CaH_2 220-5, chloroplatinate, sinters $55-70^\circ$. IV is apparently a mixt. VII, reacts neutral to litmus in H_2O , gives in acid soln with $AgNO_3$ a slight opalescence only after some time, forms no salts, quantitatively reduced by Zn dust in boiling AcOH to the *lupaninecyanoamide*, faintly yellow oil of distinct alk reaction. Chloroaurate of VIII, m 153° , picrate, m $93-4^\circ$, chloroplatinate, red, decomps 225° . Bz deriv, m 195° , only slightly sol in H_2O , with neutral reaction. Chloroaurate, m 209° . Quaternary methiodide, m $277-8^\circ$, gave on extn from alk soln with Et_4O *N*-methylupanine, viscous, strongly alk oil (chloroaurate, m 140° (decompu)), which on methylation with MeI in abs alc and distn of the quaternary NH_4 base gave a thick oil b_p $120-40^\circ$ and immediately decolorizing $KMnO_4$ in H_2SO_4 . Similar treatment of the des *N*-methylupanine gave 15% of the calcd amt of NMe_3 (as the chloroaurate), methylation of the product and extn of the neutral soln with Et_4O gave a strongly unsatd, yellowish sirup of neutral reaction in H_2O .



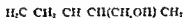
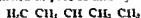
(V)

(VI)

C A R.

Lupine alkaloids. III. GEORGE R CLEMO, RICHARD RAFFER AND CHARLES R S TENNISWOOD *J Chem Soc* 1931, 429-37, cf C A 24, 123 —Clemo and Leitch (C A 22, 3665) showed that *dl* lupanine (I) is reduced to an inactive base, probably *dl*-sparteine, the work is now repeated with the active bases. Attempts to resolve I through the thiocyanates (Davies, *Arch Pharm* 235, 211(1897)) failed, however, the active forms are readily obtained by the use of the camphorsulfonates (II). *d*-I-II, in 24.2 g yield from 50 g I and 46 g II after 3 crystals from Me_2CO m $108-10^\circ$, $[\alpha]_D$ 45.5° , a satd soln of KI gives the *HI* salt, m 189° , $[\alpha]_D$ 45.5° (H_2O), of *d* I, m 40° , the thiocyanate m 184° , $[\alpha]_D$ 55.6° (1.062% H_2O soln). The mother liquor from the *d* I-II was concd, treated with KI and then with *l*-II, giving 16.1 g of *l*-I-II, m $110-3^\circ$, $[\alpha]_D$ -45.3° , the *HI* salt m 190° , $[\alpha]_D$ -43.6° , of *l* I, b, $186-8^\circ$, $[\alpha]_D$ -61.0° (3.146% Me_2CO soln), the thiocyanate m $183-5^\circ$, $[\alpha]_D$ -55.3° (1.25%

aq soln.) Equal wts. of the 2 bases, crystd from Me_2CO , give I, m 98° , d 1, Hl and red 1', heated 36 hrs at $220-30^\circ$, give I-sparteine, b₁ $130-5^\circ$, [a]_D²⁰ -11.3° (2.73% abs EtOH soln.), while this is about 6° less than the recorded value for sparteine, the picate, m $205-6^\circ$, and the Hl salt, m 231° , do not depress the m p of authentic products. Similarly, II gives d-sparteine, b₁ $133-5^\circ$, [a]_D²⁰ 15.9° , picate, m $205-6^\circ$, Hl salt, m 220° . Lupinine and PBr_3 in C_6H_6 give ω -bromolupinine (III), whose methiodide m $220-5^\circ$ (decompn) III and 20% MeOH KOH , refluxed on the H_2O bath for 16 hrs., give a mixt of δ -anhydrolupinine (IV) and lupinine Me ether, whose methiodide m $177-8^\circ$ and picate, bright yellow, m 212° . The action of heat on ω -lupinyltrimethylammonium chloride gives a mixt of IV and ω -dimethylamino-lupinine, b₁ 95° , [a]_D²⁰ -37.7° , whose dimethiodide m 306° . Karrer and Vogt (C 1 25, 900) state that they obtained in this way 30% of the anhydro base. Chlorolupinine and K phthalimide heated with a trace of Cu 72 hrs at $210-20^\circ$ give N(11)-lupinylphthalimide, m 103° , unchanged in concd H_2SO_4 after 24 hrs. electrolytic reduction in 20% H_2SO_4 gives N-lupinyl-2-anhydrosuccinimide m 55° , hydrolysis gives 11-aminolupinine (V) b₁ 98° , rapidly absorbs CO_2 from the air and yields a Br deriv m $131-2^\circ$ V and $(\text{CH}_3\text{CO}_2\text{Me})$, heated to 160° for 5 min and then slowly to 100° during 1 hr, give a mixt of dilupinylsuccinimide, m $225-6^\circ$ and lupinylsuccinimide, m 137° (methiodide, m 200°) electrolytic reduction of the amide gives N(11)-lupinylpyrrolidine, b₁ 153° , whose methiodide m 302° (decompn). The amide is not reduced under these conditions. $\text{CH}_3(\text{CH}_2\text{CO}_2\text{Me})$ and V give Me N(11)-lupinylglutaramide, m $73-6^\circ$ and dilupinylglutaramide, m $193-5^\circ$. It has not been possible to convert the Me ester into the glutaramide and obtain compds. isomeric or identical with lupanine. α -Methylloxysparteine and Br in C_6H_6 give a Br deriv, m 142° . HBr salt, m 214° , the Br is very firmly held as the compd is not attacked by EtOH KOH, HBr, KOH at 180° or boiling $\text{C}_6\text{H}_5\text{N}$. IV. Synthesis of octahydropyridocoline. G R CLEMO AND GEORGE ROWNTREE RAMAGE Ibid 437-42—Because of the work of Winterfeld and Holtschneider (preceding abstrs.), the following report is now offered. Distn of lupinic acid with soda lime gives a basic mixt from which, after catalytic reduction, an inactive base (I), $\text{C}_{11}\text{H}_{17}\text{N}$, can be readily isolated, this b₁ $43-5^\circ$, and has a strong non terpene-like basic odor, the picate, bright yellow, m $193-4^\circ$, methiodide, m $333-5^\circ$ (decompn) chloroacetate, light yellow, m $116-7^\circ$. 1-piperidine-2-carboxylate and $\text{Br}(\text{CH}_2)_3\text{CN}$ with K_2CO_3 give 90% of γ -2-carboxypiperidinobutyronitrile, b₁ 170° (an excess of nitrile gives a quaternary salt, m 184°) hydrolysis with EtOH-HCl gives 73% of Et γ -2-carboxypiperidinobutyrate, b₁ 119° . Na in PhMe gives 50% of Et 1-ethyl-2-hydropyridocoline-2-carboxylate, orange oil, b₁ 153° , heating with 20% H_2SO_4 on the H_2O bath for 3 hrs gives 70% of 1-ethyl-2-hydropyridocoline, yellow oil, b₁ 107° (semicarbazone, m 215° , methiodide m 210°). Zn and concd HCl reduce this to octahydropyridocoline (II), b₁ $43-5^\circ$, b₂ 73° , picate, yellow, m 213° (decompn), methiodide, m 223° , chloroacetate, old gold m 170° . Reduction of the keto deriv with Na Hg and abs. EtOH gives the 1-HO deriv of II, b₁ 120° , m $65-8^\circ$. Since I is not II or 3-ethylpiperidocoline, it follows that either lupinine is not represented by III or structural changes have occurred in the production of I. Such possible changes were mentioned in part II and by W and H.



Therein, a crystalline glucoside from the seeds of *Thereticia nerifolia*. L P RAMASWAMI AYYAR Proc 15th Indian Sci Cong 1923, 161—The defatted seed kernels of the yellow oleander yield on extr with alc about 1.5% of a glucoside which has been obtained in the form of glistening white plates, m 215° after softening at 190° . Analysis gave C 56.2, H 7.2, O 36.6 OH 2.4, MeO 3.0, [a]_D²⁰ -39.2° . Hydrolysis with 1% H_2SO_4 gave 57% sugar ([a]_D²⁰ -22.7° osazone, m 205° , pentaacetate, m 180°) and 48% nonsugar. From the nonsugar, an acid was obtained, m 150° .

Recent investigations on the constitution of pectins. MARC BRIDEL. J. Pharm chim [8], 13, 99-131 (1931)—A detailed review with structural formulae, a summary and 39 references are given of the investigations of Lblich and Schubert, et al., on the pectins of the sugar beet, fruits and of flaxseed (cf C 1 20, 2319, 3310, 22, 906, 23, 1623, 3465, 24, 65).

Ring synthesis of 3-methyl-4-acetylpyrrole-5-carboxylic acid. HANS FISCHER,

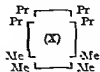
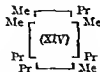
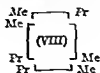
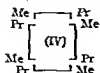
HEINRICH BEYER AND FRIEDRICH ZAUKE. *Ann* 486, 55-70 (1931); cf. C. A. 22, 2509 — In the previous synthesis of opyropyrrole from $\text{H}_2\text{NC}_4\text{H}_4\text{COMe}$ and $\text{MeCOC}_4\text{H}_4\text{COC}_4\text{H}_4\text{H}$ there was obtained a by product, m. 251°, removal of CO_2 gave a pyrrole, m. 83°, which is shown by synthesis to be 2-methyl-5-acetylpyrrole (I). 2-Methyl-5-acetylpyrrole-4-carboxylic acid (II), m. 305° (decompn.), gives an *Et* ester (III), m. 123°. II gives a hydrazone, needles, which do not m. 300°; phenylhydrazone, golden yellow, m. 238°, in AcOH PhNHNH_2 gives a pyridazine deriv., decompn. 324°, oxime, decompn. 244° boiling in NaOH gives the oxime anhydride, m. 233° (decompn.), heating the latter with N_2H_4 and EtONa in a bomb tube gives the ketazine of II. The oxime gives a *Br* deriv. in AcOH at 40°. III is reduced by AlH_3 in Et_2O , giving the carbinol, $\text{C}_{10}\text{H}_{15}\text{ON}$, m. 142°. III gives a *p*-dimethylaminobenzal deriv., cinnabar red, m. 231°; the acid soln. has a strong blue green fluorescence. Decarboxylation of II or the action of EtMgBr upon 2-methylpyrrole gives I, m. 80°, phenylhydrazone, oxime, m. 150° (*Ac* deriv. m. 147°), ketazine, $\text{C}_{11}\text{H}_{15}\text{N}_4$, yellow, m. 183°, semicarbazone, m. 200°, benzylidene deriv., yellow, m. 197°. I with 2 mols. *Br* yields a *tri Br* deriv., decompn. not sharp, 3 mols. *Br* give a *di Br* deriv., m. 161°. 2-Methyl-5-ethylpyrrole and $\text{MeC}_4\text{H}_4\text{N}_2\text{Cl}$ give an *azo* deriv., $\text{C}_{11}\text{H}_{15}\text{N}_2\text{Cl}$, orange red, decompn. 178°. PhN_2Cl gives a yellow dye, $\text{C}_{11}\text{H}_{15}\text{N}_2\text{Cl}$, decompn. 174°, the dye with $\text{HO}_2\text{SC}_4\text{H}_4\text{N}_2\text{Cl}$ m. 240-50°. III and 2 mols. *Br* give a *di Br* deriv., m. 145°, the *tri Br* deriv. m. 165°. The isooopyrrole obtained by reduction of I, heated with 2 mols. NH_4OH HCl and 1 mol. Na_2CO_3 7 hrs., gives 2,5-heptacondione dioxime, m. 150°. The reaction product of 2-methyl-3-carbethoxypyrrole (IV) and EtMgBr , treated with AcBr , gives 2-methyl-3-carbethoxy-5-acetylpyrrole (V), m. 152°, phenylhydrazone, m. 144°, oxime, m. 191°, hydrazone, m. 116°, *di Br* deriv., m. 214°. 2-Ethylpyrrole (VI) and EtMgBr , heated 4 hrs. and then treated with HCO_2Et , give triethylpyrrolylmethane, m. 162° (cor.), heating with I OH for 2 hrs. gives 2-ethyl-5-formylpyrrole, m. 52°. 2-Methyl-5-formylpyrrole, m. 70°, oxime, m. 153°. VI, MeMgI and ClC(OMe)_2 give 2-ethyl-5-carbethoxypyrrole, b.p. 134°, m. 48°, heating with said NH_4OH 12 hrs. at 150-65° gives α -ethylpyrrole- α' -carboxylic amide, m. 112°. IV and AcCl with AlCl_3 and CS_2 give V and a deep red compd., decompn. 194°. α -Acetylpyrrole and β - $\text{Me}_2\text{NC}_4\text{H}_4\text{ClHO}$ give the chalcone deriv., $\text{C}_{11}\text{H}_{15}\text{ON}_2$, orange, m. 200° (cor.)

C. J. WEST

Synthesis of some pyrroles and dipyrrolylethanones. H. FISCHER AND PIERRE VIARD. *Ber* 64B, 193-200 (1931) — The position of one or more CO_2H groups in the chlorophyll and uroporphyrin mols. is still in question. These CO_2H groups are readily split off and therefore cannot be present as LiCO_2H residues. One possibility is that they are attached to the methine groups joining the pyrrole nuclei and it was therefore undertaken to synthesize pyromethenes carboxylated on the methine group by condensing pyrroles having an α - CCl_3CO side chain with trisubstituted pyrroles and then saponify the CCl_3 group. 2,4-Dimethylpyrrole (I) with CCl_3CN yielded the 5-trichloroacetyl deriv. (II), readily obtained by hydrolysis of the imine and reduced by Zn dust AcOH to the 5-*Ac* deriv., thus establishing its structure. Similarly, I with NCCO_2Et gave *Et* 2,4-dimethylpyrrole-5-glyoxylate (III), the free acid (IV) is stable toward *Br*. Cryptopyrrole likewise yields the 5-trichloroacetyl deriv. (V), reduced to the 5-*Ac* deriv. (VI) which is readily obtained from cryptopyrrole with glacial AcOH . With alc. KOH , the imine- HCl of V splits off CHCl_3 to give α -cyanocryptopyrrole. 2,3,4-Trimethyl- and 2,4-dimethyl-3-acetylpyrroles also yielded 5- CCl_3CO derivs. Unfortunately, the condensation of these CCl_3CO derivs. with di- and trisubstituted pyrroles has as yet given no well-defined products, there result dyes with no characteristic absorption and having no tendency to crystallize. In phylloporphyrin it is not yet definitely established whether the 2 pyrrole nuclei are united through a group $-\text{CH}_2\text{CH}_2-$ or $-\text{CMe}-$. To throw light on this point bis(2,4-dimethyl-3-acetyl)pyrrolylmethane (VII) and the *Pr* homolog (VIII) were prepd. by condensation of 2,4-dimethyl-3-acetylpyrrole with the appropriate aldehydes, but dehydrogenation to the methenes has thus far yielded no cryst. products. As the pyrrole nuclei in chlorophyll may possibly be united through a $-\text{CH}_2\text{CO}-$ group, a ClCH_2CO group was introduced into I, 2,3,4-trimethylpyrrole and cryptopyrrolecarboxylic acid and the products were treated with the Grignard compds. of di- and trisubstituted pyrroles. The resulting dipyrrolylethanones with α - α' -brominated pyromethenes gave dyes which have not yet been obtained in cryst. form. II (22 g from 10 cc. I), m. 108-9°, is stable toward boiling alc. 2,4-Dimethyl-5-carbethoxy-3-trichloroacetylpyrrole (60-70% yield), m. 173-4°. 3-*Br* deriv. of II (11.5 g from 10 g I with *Br* in AcOH), m. 145-6°. V (4 g from 2 cc. cryptopyrrole), m. 101-2°. VI (70% from V), m. 111-2°. 2,3,4-Trimethyl-5-trichloroacetylpyrrole (90%), m. 114-5°. 2,4-Dimethyl-3-acetyl-5-trichloro-

acetylpyrrole, m. 145-6°. 2,4-Dimethyl-3-bromo-5-chloroacetylpyrrole, m. 184-6°. III, m. 60-1°. IV, m. 164-5° (decompn). VII (87%), m. 208-9°. VIII, m. 213-4°. MeCH ClCHCO gives a product m. 236-8° whose compn, however, does not agree entirely with that calcd for the propenyl analog. of VIII. (3,4,5-Trimethylpyrrol-3',5'-dimethylpyrrol)-2,2'- α ethanone, m. 168°. 3,5-Dimethyl-3',5'-dimethyl-4'-ethyl analog, m. 162-3° (in 1 prepn there was obtained as by-product an isomeric (probably dimeric) product, m. 229° (cor)). 3,5-Dimethyl-4-ethyl-3',5'-dimethyl compd, m. 156°. 3,4,5-Trimethyl-3',5'-dimethyl-4'-ethyl compd, m. 157°. Bis(3,5-dimethyl) compd, m. 179° (cor). 2,4-Dimethyl-1-ethyl-5-chloroacetylpyrrole with 1 mol SO_2Cl_2 in Et_2O gives a mono-Cl deriv, m. 123-5°, whereas with 2 mols SO_2Cl_2 is obtained a product which, after washing with ice H_2O , is 3-ethyl-4-methyl-5-chloroacetylpyrrole-2-aldehyde, m. 113-4° (Chloroacetyl)cryptopyrrolecarboxylic acid, obtained with ClCH_2CN , m. 212-3°. 2,3,4-Trimethyl-5-arsinoacetylpyrrole, from the 5- ClCH_2CO deriv and PhNH_2 , m. 194-6°. 2,4-Dimethyl compd, m. 207°. 2,4-Dimethyl-3-bromo compd, m. 205-10° (decompn). 2,4-Dimethyl-3-ethyl compd, m. 181-3°. C. A. R.

Porphyrin syntheses. XXXVII. Synthesis of tetramethyltetrapropylporphyrins I-IV and of octapropylporphyrin. H FISCHER, MARTIN GOLDSCHMIDT AND WILHELM NÖSSLER. Ann. 486, 1-54(1931), cf C A 25, 1236 —Although the 4 possible isomers of etioporphyrin (tetramethyltetraethylporphyrin) have been synthesized, their high m ps. make a detn of their differences difficult, for this reason the corresponding tetramethyl tetrapropylporphyrins have been prepd. 2,4-Dimethyl-5-carbethoxypyrrrole and EtCOCl with AlCl_3 give 83% of the 3-propionyl deriv (I), m. 140°. in AcOH Br and I give the 2-bromomethyl deriv, m. 145° (cor). Heating 100 g I, 55.5 g Na, 667 cc. abs. EtOH and 50 cc. 77% $\text{N}_2\text{H}_4 \cdot \text{H}_2\text{O}$ 12 hrs. at 160-5° (25-35 atm) gives 25-6 g 2,4-dimethyl-3-propylpyrrole (II), b.p. 98-100°, m. 13.5°, d^{20}_4 0.8888, n_D^{20} 1.49243, picrate, yellow, m. 132° (cor). $\text{HO}_3\text{SC}_6\text{H}_4\text{N}_2\text{Cl}$ gives an azo dye, $\text{C}_{18}\text{H}_{18}\text{N}_4\text{O}_6\text{S}$, brown red, m. 248°. II, (18 g) HCN and HCl in $\text{CHCl}_3\text{-Et}_2\text{O}$ give 14 g 2,4-dimethyl-3-propyl-5-formylpyrrole, m. 105°, NCCO_2Et gives the 5-Et glyoxylate deriv, m. 192°. II in 30% H_2SO_4 with CrO_2 gives methylpropylmaleimide (III), m. 57°. II gives an oxime, m. 198°. The residue from II is the pyrazoline of I, b.p. 141-5° (yield, 45 g), di-HCl salt, m. 145° (cor); dipicrate, deep yellow, m. 137° (cor). 2,4-Dimethyl-3-propionylpyrrole and 77% $\text{N}_2\text{H}_4 \cdot \text{H}_2\text{O}$ in abs. EtOH , boiled 5 hrs., give the ketazine, canary-yellow, m. 159° (cor.). II and Br in AcOH-CCl_4 give the HBr salt, blue-violet leaflets, m. 130° (HCl salt, brown red, m. 178°, picrate, yellow-brown, m. 182°), of 5-bromo-5'-bromomethyl-4,3'-dimethyl-3,4'-dipropylpyrromethene, brown red, decompn. 207°. The methene with HBr-AcOH , heated 4 hrs. at 170° or better with $(\text{CH}_3\text{CO}_2\text{H})$, 10 min. at 180° gives 10-15% of the tetramethyltetrapropylporphyrin (IV), m. 200°; Cu complex, needles, m. 295°; Mg complex, red leaflets, m. 249°; Fe complex, dark leaflets, m. 338° (cor), the bromohemin m. 333°, the iodohemin m. 310°; the nitroporphyrin, brown red, m. 272°. Shaking IV in $\text{CHCl}_3\text{-AcOH}$ with PbO_2 until the porphyrin spectrum has disappeared, gives aurothoporphyrinogen, yellow, crystals with 2 mols H_2O , m. 262°; reduction with Na-Hg again gives IV; photooxidation gives III. The reaction product of II and EtMgBr , treated with ClCO_2Et , gives 70% of 2,4-dimethyl-3-propyl-5-carbethoxypyrrrole, m. 98° (cor); the 2-bromomethyl deriv, m. 148°; the crude Br deriv, heated with (HCHO) , in MeOH 3-4 hrs., gives 4,4'-dimethyl-3,3'-dipropyl-5,5'-dicarboethoxypyrrromethane (V), m. 166°; the free acid is not stable and could not be recrystd; the crude product decompn. about 140°, boiling in AcOH for 4 hrs. gives some 2-methyl-4-propylpyrrole and a small amt. of crystals, m. 121-2°. The dry acid



and Br in $\text{HCO}_2\text{H-AcOH}$ give 5,5'-dibromo-4,4'-dimethyl-3,3'-dipropylpyrromethene (VI), orange-brown, m. 146°, as the HBr salt, dark violet, does not decomp. at 270°;

Coloring matter of awobana. CHIKA KURODA *Proc Imp Acad Tokyo* 7, 61-3 (1931).—The flower of "Tsuyukusa" (*Commelina communis*), which is azure blue, is used in the prepn of awobana paper, which was used in these expts. The dye is insol in MeOH but is sol in H₂O, from which the dye is pptd by MeOH as an ultramarine-colored powder. The ash contains (14.5-15%) principally K and Mg, with traces of Fe, Al, Ca, PO₄ and SO₄ are the important acid radicals. Heating with MgO gives about 1% NH₃. The coloring matter, acidified, turns purple to red without losing its sol in H₂O. Hydrolysis gives about 31% of sugar, which appears to be 1 mol of a monosaccharose (glucose) the red aglycon is sol in EtOH but insol in other org solvents. Decompn with alkali gives *p*-coumaric acid and *p*-HOC₆H₄Ac, the acid is obtained even with cold 1% aq alkali while the ketone is formed only after heating
C J WFS

The formation of films of drying oils (KAPPELMEIER) 26. Electric moment and molecular structure III Double and triple bonds and polarity in aromatic hydrocarbons (SMYTH, DORRIS) 2. The birefringence of saffrole (L'AUTHENIER, BART) 2. X ray examination of the crystal structure of resorcinol (SARKAR) 2. Crystals of 3,4,3',4',6'-pentamethoxydiphenylmethane 2-carboxylic acid (LENGYEL) 2. Ultra violet absorption spectra of the nitriles and amides of 2 methyl-3 pentene (CASTILLE, RUTPOL) 3. Melting point curves of the monobasic fatty acids (KING, GARNER) 2. The decomposition of hydrocarbons in the positive ray tube (STEWART, OLSON) 3. The chemical action of ultra violet light on the alkyl iodides (FISCHWILLER) 3. Significance of the structure of the hydrocarbon residue on the velocity and equilibrium position in organic reactions (HEROLD, WOLF) 3. The absorption of aqueous solutions of tartaric acid (BRUHAT) 3. Electrical properties of molecules (ATEN) 2. Interaction of epichlorohydrin and cyclohexene oxide with alkali and NH₃ halides (SEY, *et al*) 2. Recovery of solvents such as those used in the manufacture of the alkali metal phenyl glyoxines (U S pat 1,798,713) 13.

Beilsteins Handbuch der organischen Chemie. 4th ed. Erstes Erg.-Werk, d. Literatur von 1910-19 umfassend. Bd. VI. Als Erg. d. 6 Bds. d. Hauptwerkes (System-Nr 499 bis 608). Issued by Deutschen Chemischen Gesellschaft. Edited by FRIEDRICH RICHTER. Berlin. J. Springer. 642 pp. Linen, M 128.

MACBETH, A. KILLEN. Organic Chemistry for Medical Intermediate Science and Pharmaceutical Students. London. Longmans Green and Co. 296 pp. 6s. 6d.

VLASSOPOLLOS, VLASSOS. Über die sterische Hinderung bei Reaktionen von Aminosäuren und Polypeptiden, zugleich ein Beitrag zum Wesen der sterischen Hinderung. Leipzig. Gustav Fock G m b H. 30 pp.

Organic oxygen compounds. HENRY DREYFUS. Fr 697,726, June 20, 1930. O compds of C are hydrogenated in the presence of catalysts composed of or contg Fe or Co in the form of compds having the metals in the acid radical, *e g.* ferrites, ferrates, cobaltites or cobaltates of alkali or alk earth metals. Examples are given of the prepn of EtOH and higher alcs along with aliphatic acids and aldehydes from CO and H₂ and from a mixt of MeOH, CO, CO₂ and H₂. Fr 697,727 describes the prepn of similar compds. in which catalysts composed of or contg compds of metals of groups 1 and 2 are used with oxy acids of elements of groups 2, 3, 4 and 5, *e g.* borates, aluminates, zincates, silicates or phosphates of alkali, or alk earth metals or Cu. Cf C A. 25, 963 and following abstrs.

Organic oxygen compounds. HENRY DREYFUS. Fr 697,896, June 25, 1930. EtOH and higher aliphatic alcs are obtained from mixts of H and O compds of C in the presence of a catalyst composed of or contg Fe, Ni or Co in the form of oxy acids of groups 2, 3, 4 or 5, or as free metals (or their compds.) in intimate assocn with other compds of oxy acids. The metals may be used in the form of borates, aluminates, zincates, silicates, phosphates or other salts of oxy acids of P.

Organic oxygen compounds. HENRY DREYFUS. Fr 698,094, June 26, 1930. Org O compds are prepd by the reaction of steam with aliphatic hydrocarbons such as CH₄, C₂H₆, C₃H₈, or C₄H₁₀. Alk earth metals or metals of groups 4, 5, 6, 7 and 8, Al, Mg or Zn or their compds may be used as catalysts. The reaction is carried out under pressure and preferably at 250-500°. Examples are given of the prepn of MeOH and higher alcs along with aldehydes, acids and ketones.

Oxidizing paraffins, etc. I. G. FARBENTIN, A-G (Manfred Dunkel, inventor). Ger 522,361, Mar. 19, 1927. Paraffin hydrocarbons, waxes, etc., are oxidized with

gases contg O in the presence of both a metalliferous catalyst and a small quantity of an org. base. Thus, paraffin may be oxidized at 170° with air in the presence of Mn soap 0.5 and $(\text{CH}_3)_2\text{N}_2$ 0.3%.

Higher olefins and diolefins from lower olefins. PAUL FEILEA (to I G Farbenind. A.-G.) U. S. 1,799,787, April 7. In the production of compds. such as diolefins from lower olefins such as C_3H_6 , the initial gases are exposed, at a higher rate of flow than that required for the production of C_3H_6 , to the action of high temp. elec. discharges and the process is carried out in a closed cycle with continuous removal of the higher olefins and diolefins by cooling. Cf. C. A. 25, 2473.

Diolefins. I. G. FARBENIND. A.-G. Fr. 698,426, July 5, 1930. 1,3 Butylene glycol or other compds. hydroxylated in the 1 and 3 positions are dehydrated in the presence of catalysts and under pressure. Suitable catalysts are H_3PO_4 , red P, NaH_2PO_4 , or alum. Examples are given.

Cyclic aldehydes. I. G. FARBENIND. A.-G. (Georg Kalischer and Karl Keller, inventors) Ger. 519,806, Mar. 22, 1928. Addn. to 514,415 (C. A. 25, 1536). The aldehyde group is introduced into oxygenated cyclic compds. by treating these with formamide in the presence of AlCl_3 or a chloride or oxychloride of P or S. A large excess of formamide should be used. Thus, a mixt. of β naphthol and formamide may be heated to 95°, mixed with AlCl_3 , the mixt. heated to 130-40°, and the cooled product stirred with water, then heated with AcOH and filtered, β naphthol 1 aldehyde is obtained from the filtrate. Other examples are given also.

Ketones. SCHERING-KAHLBAUM A.-G. Fr. 698,230, June 30, 1930. Vapors of aromatic dicarboxylic acids or their anhydrides and vapors of aliphatic carboxylic acids are passed together over suitable catalysts. Thus, a mixt. of acetone and acetophenone is obtained by passing a mixt. of phthalic anhydride and AcOH over MnO_2 at 350-400°.

Catalytic synthesis of amines from alcohols and ammonia. HERAUX R. ARNOLD (to E. I. du Pont de Nemours & Co.) U. S. 1,799,722, April 7. In processes such as effecting reactions between MeOH and NH_3 in the presence of a catalyst such as alumina gel, the mol. ratio of alc. to NH_3 is the most important factor influencing yield and proportion of primary, secondary and tertiary amines in the product. As this ratio is increased the total conversion increases. At low ratios, e. g., 1 to 4, primary amine is the major product, while at a ratio of about 3 to 1 primary amine disappears, the product consisting of secondary and tertiary amines tertiary being formed in the larger amts. Examples with details of procedure are given.

Aromatic amines. I. G. FARBENIND. A.-G. (Johus Laux, inventor). Ger. 516,999, Feb. 16, 1927. Addn. to 515,758 (C. A. 25, 2437). The method of 515,758, for producing aromatic amines by the reduction of nitro compds. by Fe in the presence of aq. Al salt soln. is modified by replacing the Al salts by salts of other ter- or quadri-valent metals. Thus, PhNO_2 is reduced by Fe filings in the presence of CeCl_3 to produce PhNH_2 .

Monocyclic lactones. LEOPOLD RUTICKA (to Soc. anon. M. Naef & Co.). U. S. 1,799,536, April 7. See Fr. 657,971 (C. A. 23, 4483). For prep. monocyclic lactones having 14 to 18 ring members, monocyclic ketones having 13 to 17 ring members are heated with persulfuric acid. Examples with details of procedure are given.

Keto acid esters. PAUL HALBIG, FELIX KAUFER and H. PETER SCHMITZ (to A. Wacker Ges. für Elektrochemische Industrie G. m. b. H.). U. S. 1,798,937, March 31. See Fr. 654,413 (C. A. 23, 3714).

Diazo solutions. KARL SCHNITZPAHN (to General Aniline Works). U. S. 1,799,068, March 31. A solid aryl diazonium fluoborate of the benzene series such as *p*-nitrophenyldiazonium fluoborate is dissolved in an aq. soln. of salts of metals such as KCl or AlCl_3 , the metal of which has a greater affinity for the hydrofluoboric acid than for the acid residue of the salt, so that difficultly sol. fluoboric metal salts may be sepd. and concd. diazo solns. obtained.

Stable diazo compounds. SOC. ANON. POUR L'IND. CHIM. A. BALR. Fr. 697,425, June 14, 1930. Solns. of diazo compds. are treated with aromatic sulfonic acids contg more than one SO_3H group in the presence of salts of metals of the 2nd group of the periodic system, the hydroxides of which are strong bases, e. g., CaCl_2 , MgCl_2 or MgSO_4 . Cf. C. A. 24, 1391.

Hydroxycarbazoles. I. G. FARBENIND. A.-G. Fr. 698,148, June 27, 1930. The prep. of 2- or 3-hydroxycarbazole is described by treating 2- or 3-alkoxytetrahydrocarbazole with PhO at a temp. of 400-600°.

Phenols. F. RASCHIG G. m. b. H. Fr. 698,341, July 3, 1930. PhOH and its homologs are prep. by treating PhCl and its homologs with steam at a high temp. in the presence of Al_2O_3 or $\text{Al}(\text{OH})_3$ free from Fe. The catalyst may be prep. by

neutralizing a soln of $\text{Al}(\text{OH})_3$ in an alk. lye with an acid soln of a Cu salt. Instead of Al other metals of the 1st to the 4th group may be used.

Stabilizing pheools. I G FARBEINIG A-G. Ger. 522,391, Feb 24, 1929. The darkening and ultimate resinification of pheools or their solns under the influence of light and air is hindered by addn of 0.5–3% of oxalic acid or its acid or neutral salts or esters.

Alkyl substituted pheools. Soc. REICHHOLD, FLUGGER & BOECKING. Fr. 697,711, June 20, 1930. Unsaturated hydrocarbons are caused to react on pheools in the presence of metallic chlorides such as AlCl_3 , ZnCl_2 and FeCl_3 . The reaction is started by small quantities of alkyl halides or halogen acids. Thus, gaseous isobutylene is bubbled through a soln of PhOH in CCl_4 contg AlCl_3 , a small quantity of *tert*-BuCl being added at the start. Other examples are given.

Isoalkylene pheools. WALTER SCHÖLLER, HANS JOHANN and REINHARD CLERC (to Schering-Kahlbaum A-G). U.S. 1,798,813, March 31. 3-Methyl-6-isopropylene-phenol is prepd by heating 4,4'-dimethyl-2,2'-dihydroxydiphenylmethane above the m.p. and sepg the decompn products by fractional distn.

Polymerization products. I G FARBEINIG A-G. Fr. 697,437, June 14, 1930. Acid polymers or their derivs are prepd by polymerizing, in known manner, the halides or nitriles of monomeric acids capable of being polymerized and afterward substituting the halogen or CN by other radicals by appropriate treatment. By treatment with water, steam, aq solns of alkalis or moist air the acids themselves are obtained, and with less than the equiv amt of water the anhydrides are obtained. Examples are given of the treatment of polymerized acrylic acid chloride.

Condensation products. OSCAR ADLER and RUDOLF ADLER. Fr. 697,881, June 25, 1930. Salts of phenylquinolinecarboxylic acid and pyrazolones or alkylated aminopyrazolones are heated in the presence of water to a temp necessary to obtain a clear fusion. Thus, Ca 2-phenyl-4-quinolinecarboxylate is ground with 1 phenyl-2,3-dimethyl-4-dimethylamino-5-pyrazolone and a little water and heated. The mass obtained is cooled and ground.

Alkylene derivatives. I G FARBEINIG A-G. Fr. 697,786, June 23, 1930. Alkylene oxides with org compds contg OH or COOH groups or water are passed through heated pressure towers and afterward distd. Thus, ethylene glycol is prepd from a mixt. of ethylene oxide and water, and glycol monoacetate from ethylene oxide and AcOH. Other examples are given.

Derivatives of anthraquinone. I. G. FARBEINIG A-G. (Wilhelm Müller, inventor). Ger. 516,997, June 30, 1928. Derivs of anthrapyrimidine character are obtained by the action of NH_3 on anthraquinone-2,1-oxazole. Thus, C-phenylanthraquinone-2,1-oxazole is heated to 150° with aq NH_3 for 12 hrs in an autoclave to give 2-amino-C-phenyl-1,9-anthrapyrimidine, m. 343–5°. Further examples are given.

Benzimidazole derivatives. I G FARBEINIG A-G. (Wilhelm Eckert and Otto Braunsdorf, inventors). Ger. 522,173, June 19, 1928. 4-Dialkylacetylnaphthylenebenzimidazole-5-carboxylic acids are prepd by heating naphthylenebenzimidazole-*peri*-dialkylindaodiones with KOH. The initial materials are prepd by oxidizing acenaphthenedialkyl-*peri*-indaodiones (cf C.A. 4, 2144), and condensing the resulting naphthalene-4,5-dialkylindaodione-1,8-dicarboxylic acids or their anhydrides with α -diamines. Examples are given.

Benzothiazole derivatives. I. G. FARBEINIG A-G. (Heinrich Jensch, inventor). Ger. 522,059, Sept 18, 1928. See Fr. 680,384 (C.A. 24, 3801).

Guanidine derivatives. R. FRIEDRICH VON GODIN. Ger. 522,057, July 10, 1929. Salts of alkylene diguanidines are prepd by fusing higher alkylene diamines with guanidine salts. The reaction proceeds with evolution of NH_3 . Examples are given. Cf C.A. 23, 1649.

Hydroxydiphenylamine derivatives. I G FARBEINIG A-G. Fr. 697,764, June 21, 1930. Derivs of 3-hydroxydiphenylamine are prepd by heating 1-methyl-2-amino-4-hydroxybenzene with aromatic amines contg no sulfonic, carboxylic or nitric groups in the presence of HCl or agents liberating HCl. Examples are given of the prepn of the 6-methyl-4,6-dimethyl and 4-chloro derivs.

Piperazine derivatives. Soc. DES USINES CHIM. RHÔNE-POULENC. Fr. 698,687, Oct 7, 1929. Mono- and di-substituted derivs of piperazine are prepd by the reaction of ethylene oxides on piperazine hydrate, in excess to obtain the monosubstituted, and out in excess to obtain the disubstituted derivs. Examples are given of the prepn. of piperazinedimethylethylcarbinol, m. 77–78°, piperazinebis(dimethylethylcarbinol), m. 230° (with decompn), *N*-benzoylpiperazinedimethylethylcarbinol, m. 33°, piperazinebutyloxymethylmethylcarbinol (picrate, m. 204°), piperazinemethylbenzylcarbinol,

m 63-64°, piperazinebis(methylbenzylcarbinol), m 175°, piperazinemethylphenoxymethylcarbinol, m 58-59°, piperazinebis(methylphenoxymethylcarbinol), piperazine- α quinoline (picrate, m 244°), and quinolinepiperazinemethylcarbinol

Pyridine derivatives. SCHERING-KAUBAUM A-G (Curt R  th, inventor) Ger 522 060, April 15, 1927 Derivs of 2-chloropyridine substituted in the 5-position or the 3,5-positions with NO_2 , COOH , or halogen, are prepd by halogenating the correspondingly substituted *N*-alkyl- or *N*-aryl 2-ketopyridines. Thus, *N*-methyl 2-keto-5-nitropyridine, treated at 150-160° with PCl_5 contg a little POCl_3 , yields 2-chloro-5-nitropyridine. Other examples are given.

Xylenol derivatives. I G FARBERND A G. Fr 697 821, June 23, 1930 Arylides of asym- α -xylenolcarboxylic acid are prepd by heating the carboxylic acid (obtained by heating alkali salts of α -xylenol of the constitution C_6H_4 , C_6H_3 , $\text{OH} = 1:2:4$) in the presence of condensing agents with arylamines contg no COOH or SO_3H groups. Examples are given of the prepn of the anilide (m 178°), the *p*-chloroanilide (m 209-210°), the *o*-toluide (m 296-297°) and a list of constituents with the m ps of the compds is given.

Stable leuco indigo preparations. IMPERIAL CHEMICAL INDUSTRIES, LTD Fr 699 491, April 2, 1930 Dispersed indigo, the particles of which are much smaller than those of ordinary finely divided indigo (ϵ , below 10 μ) is reduced, by means of alkali and reducing sugar, one mol proportion or less of reducing sugar and 7 mol proportions or less of alkali being used for one mol proportion of indigo Cf C. A. 25, 1539

Trialkyl phosphates. WM J BANISTER (to Commercial Solvents Corp.). U. S. 1,799,349, April 7 Tributyl phosphates are obtained by the reaction of Al butoxide with P oxychloride in C_6H_6 at a temp below 15°, adding water to hydrolyze AlCl_3 and fractionally distg

Calcium tartrate. SOCI  T   INDUSTRIELLE DU LANGUEDOC (S A) Ger 516,673, April 24, 1929 Crystd Ca tartrate is freed from its water of crystn by heating with water to above 120° in an autoclave. Thus, the salt contg 72.3% Ca tartrate and 27.7% water, is mixed with water and heated to about 160° and steam under pressure led in. The resulting salt has a water content of 0.4%.

Aryl isothiocyanates. J D RIEDEL-E DE HLAEN A-G (Karl H Slotta and Helmut Dressler, inventors) Ger 522,058, Oct. 18, 1928 Aryl isothiocyanates are prepd by the reaction $\text{RNHCSNH}_2 + \text{COCl}_2 = \text{RNCS} + \text{HCl} + \text{NH}_4\text{Cl} + \text{COS}$. The reaction is effected in an anhyd solvent, e.g., C_6H_6 , while cooling. The arylidithiocarbamate required as initial material is preferably prepd with exclusion of water, e.g., by treating an aromatic amine in C_6H_6 soln with dry NiH_2 , and then with CS_2 , while still passing in NH_3 . Examples are given.

Halogen derivatives of α -chloronaphthalene. I G FARBERND A-G (Georg Rosch and Wilhelm Bauer, inventors) Ger 516 671, Oct. 23, 1928 1-Chloronaphthalene-8-sulfonic acid is halogenated. If the SO_3H group is retained in the mol after halogenation, it can be removed or replaced by halogen if desired. Several examples are given. In one, Na 1,4-dichloronaphthalene 8-sulfonate is prepd by chlorinating Na 1-chloronaphthalene 8-sulfonate. In another, 1,4,8-trichloronaphthalene, m 130°, is obtained by further chlorinating the same sulfonate as in the preceding example.

Tetramethylthiuram polysulfides. ADRIEN CAMBROY (to Roessler & Hasslacher Chem Co) U S 1,798,588, March 31 After reaction of an aq soln of NH_4Cl with formaldehyde, the soln is treated to remove the major portion of unreacted NH_4Cl and by products formed, and the resulting soln is treated with CS_2 followed by an aq soln of an alkali metal hydroxide while cool to produce an alkali metal salt of dimethyl dithiocarbamic acid, and the soln is treated with S_2Cl_2 to form tetramethylthiuram tetrasulfide.

Acrylic acid chloride. I G FARBERND A G Fr 697,311, June 30, 1930 See Brit. 333,079 (C A 25, 524)

Salts of sulfuric acid esters of nitroanthrahydroquinone. I. G FARBERND A-G (Arthur Luttringhaus, Heinrich Neresheimer and Wilhelm Schneider, inventors) Ger 516,845, Nov 6, 1927 Nitroanthraquinone is treated with metals in the presence of tertiary bases and halides or anhydride of H_2SO_4 . Thus, 1-nitroanthraquinone and Cu powder is added to pyridine and Me sulfuric acid chloride. Further examples are given.

Sulfuric esters. RICHARD M DEANESLY (to Shell Development Co) Can. 310,439, April 14, 1931 Alkyl sulfuric esters are prepd by contacting liquid olefins with H_2SO_4 at a const temp obtained by permitting the vaporization of some part of

the hydrocarbons and condensing the vapors in a condensing device apart from the reaction vessel, and returning the condensed hydrocarbons to the reaction vessel.

Silicic acid ester. TOSIYASO SAKAMI and KANSAI PAINT KAKUSHIKI KAISHA, Japan 90,227, Feb. 10, 1931. Silicic acid ester is prepd. from SiCl_4 and MeOH , EtOH , BuOH , AmOH or benzyl alcohol. The product is heated above 100° to expel HCl . Then a small amt. of containing HCl is removed by mixing with slight excess of oxides or hydroxides of Pb , Hg or Ag at about 60° . The product contains no trace of HCl . The metallic chlorides formed are insol. in H_2O or alc. and sep. by decantation.

Acyated esters of hydroxy acids. FOWARD J. POWERS (to Commercial Solvents Corp.) Can. 310,851 Apr. 28, 1931. A mixt. of a hydroxy aliphatic acid and an ester is refluxed in the presence of a catalyst. Suitable catalysts are any concd. HCl , dry HCl , anhyd. NaHSO_4 , concd. H_2SO_4 , benzene, sulfonic acids, etc.

Compounds of dialkylbarbituric acids and 1-phenyl-2,3-dimethyl-4-dimethylamino-5-pyrazolone. PETER FRIEDRICH (to Hoffmann La Roche Inc.) U. S. 1,708,570 March 31. Mol. proportions of 1-phenyl-2,3-dimethyl-4-dimethylamino-5-pyrazolone and a compd. such as isopropylalylbarbituric acid are dissolved in a common solvent such as acetone or C_6H_6 and a compd. of the 2 side chains sep. as crystals which are removed from the solvent before any excess of the starting materials crystallizes out. The products form yellow crystals sol. in org. solvents.

Products from 1-phenyl-2,3-dimethyl-4-dimethylamino-5-pyrazolone and 5,5-substituted barbituric acids. GEORG FRIEDRICH (to Schering Kuhlmann A. G.) U. S. 1,708,471, March 31. A substantially colorless solid product is obtained by addition of a molten mixt. of 1-phenyl-2,3-dimethyl-4-dimethylamino-5-pyrazolone and a 5,5-substituted barbituric acid such as 5,5-diethylbarbituric acid.

Recovery of dilute acids in esterification processes. IGNACE J. KRCHMA (to Commercial Solvents Corp.) Can. 311,002, Mar. 11, 1931. In the production of esters, e. g., butyl acetate, by the reaction of acid and alc., the unchanged acid is recovered from the aq. distillate fraction by extr. with the alc. and then the water is removed from the alc. soln. by distn.

Fatty acids. I. G. FARMINEK A. G. Pr. 647,695, June 18, 1931. Hydrocarbons are oxidized to fatty acids by means of oxides of N or HNO_3 , and the product of the oxidation is submitted to a temp. below 200° or to a reduction, preferably in the presence of catalysts or under high pressure or both, before the final distn.

Sulfonic acids. N. V. de HATAAPROEF KUNSTOFEN MAATSCHAPPIJ Pr. 699,630, June 11, 1930. Sulfonic acids, the Ca salts of which are sol. in water, are prepd. from sulfonic acids the Ca salts of which are insol. by treating the latter with concd. or fuming H_2SO_4 or an agent acting in the same way. Examples are given.

o-Xylenedicarboxylic acid. I. G. FARMINEK A. G. (Lampold Laska and Oskar Haller, inventors) Ger. 522,664, May 29, 1931. An acid believed to be 1,2-dimethyl-4-hydroxybenzene-5-carboxylic acid is prepd. by heating an alkali salt of 1,2-dimethyl-4-hydroxybenzene with CO_2 under pressure. An example is given.

Acetic acid. HERMANN HUBA, Austrian 121,048, Nov. 15, 1930. Conc'd. AcOH is recovered from dil. AcOH by exp. the superheated vapor of the dil. acid with a high-boiling solvent for AcOH to which 5-10% of pure or conc'd. MeOH has been added. Cl. C. A. 25, 2440.

Acetic acid. VERBODEN VER KUNSTOFEN EN METAALORGANISCHE PRODUCTEN Pr. 698,698, June 20, 1930. AcOH and AcOMe are made by catalytic synthesis from Me_2O and CO or gases contg. them at temps. of $250-170^\circ$ and pressure of 300-600 atm. Neutral or acid phosphates of Cr and of a metal of the 3rd group, e. g., Al are used as catalysts.

Polyvinyl alcohol. CONSORTIUM FÜR ELEKTROCHEMISCHE INDUSTRIE G. M. B. H. (Wolfram Haefliger and Willy O. Herrmann, inventors) Ger. 510,000, Sept. 23, 1928. Addn. to 450,295. Polyvinyl halides are saponified by alk. reagents in the presence of catalysts. In the example, polyvinyl chloride is saponified with NaOH in the presence of light and nitril nitrate. A 90% yield is obtained.

Methylamine. I. G. FARMINEK A. G. Pr. 698,540, July 8, 1930. The halides of dicarboxylic or sulfonic acids or their N -methylol deriva. are treated with CH_3O under heat and pressure. Examples are given of the treatment of phthalimide, α -naphthalene acid, N -methylolphthalimide and succinimide.

Naphthalene purification. GEORGE SCHREPTER (to Newport Mfg. Co.) U. S. 1,800,161, April 7. Naphthalene contg. impurities capable of producing a red coloration on treatment with conc'd. H_2SO_4 is treated, at a temp. above its m. p., with a porous material such as absorbent earth contg. small quantities of metal oxides such as lime

(d) CH_2O_4 and Fe, and from defibrinated blood by (e) putrefaction or satn. with illuminating gas, was converted into the Na or K salt and purified by crystn. from pyridine or MeOH-KOH-AcOH. Various cryst. forms were obtained depending on the procedure, and these forms were mutually convertible. A. W. DOX

Absorption spectra of the fractions of serum albumins. E. FALTIN. *Magyar Chem. Folyóirat* 37, 74-87(1934).—The ultra-violet absorption spectra of serum albumin of horse and oxen blood were examd. No distinction can be made in the different bloods because of the very varying action of different animals of the same species. The difference is probably caused by the fact that albumin and globulin are not homogeneous albumins but mixts of several proteins with different absorption spectra. The ratio of the mixt. varies probably in each individual. Gradual satn. of sera with $(\text{NH}_4)_2\text{SO}_4$ sepd. 6 protein fractions, the absorption power of which was less the more $(\text{NH}_4)_2\text{SO}_4$ was needed for their sepn. Of the various fractions the contents of tyrosine and tryptophan diminished in proportion to their absorption power. S. S. DE FINÁLY

The nature of the chromaffin reaction. P. GERARD, R. CORDIER AND L. LISON. *Bull. hist. appl. physiol. et path.* 7, 133-9(1930). *Physiol. Abstracts* 15, 331.—The red color obtained by the action of oxidizers on adrenaline is due, not to a simple oxidation product, but to an addn. compd. formed between the quinone deriv. produced in the oxidation and the unchanged adrenaline. The chromaffin reaction, then, is to be interpreted as a dehydrogenation and transformation of adrenaline into a quinone deriv., which then combines with as yet unchanged adrenaline to give a strongly colored substance. G. G.

Histochemical study of the chromo-argentaffine substance in Kultschitzky's cells. R. CORDIER AND L. LISON. *Bull. hist. appl. physiol. et path.* 7, 140-8(1930). *Physiol. Abstracts* 15, 331.—These are special cells in the intestinal epithelium. Guinea-pig duodenum was fixed in 10% neutral formal and imbedded in paraffin, and the diazo reaction for the phenol group applied to the sections. The presence of 2 OH groups was then proved by the usual oxidation tests. Thus the phenolic constitution of the argentaffine granules, previously postulated, was definitely confirmed. G. G.

The red fluorescence which certain bilirubin derivatives show in ultra-violet light. CH. DUBÉ, *Compt. rend. soc. biol.* 103, 371-4(1930). *Physiol. Abstracts* 15, 303.—Very dil. solns of bilirubin with NH_3 , $(\text{AcO})_2\text{Zn}$ and a drop of I show a striking red fluorescence with ultra-violet light. The fluorescence is intense, like that of the porphyrins. Bilirubin in CHCl_3 soln. plus Br gives a blue pigment which shows a red fluorescence. The blue color of cholecyanin comes out with soda and NaNO_2 . All these fluorescent solns. show several absorption bands. A distinctive band in the red is present even with extreme diln. It is impossible to say if all the fluorescent pigments consist of bilucyanin, or whether several complex combinations exist with I, Br and Zn. Also in *Arch. intern. pharmacodynamie* 38, 134-9(1930). G. G.

Models for the capillary electrical theory of irritability. J. KLINKE. *Arch. ges. Physiol. (Pflügers)* 227, 110-31(1931).—Mech. models (gelatin membranes in solns of electrolytes) were prepd. to illustrate the transmission of elec. impulses in the body. ARTHUR GROLLMAN

Effect of heat upon the activating efficiency of enterokinase. JAMES PAGE. *Biochem. J.* 25, 1-7(1931).—The course of the heat inactivation may be described by the unimol. expression. The optimum pH for enterokinase lies between 6 and 7. BENJAMIN HARROW

Optimum hydrogen-ion concentration and temperature of the style enzyme of *Pecten maximus*. ALASTAIR GRAHAM. *Proc. Roy. Soc. (London)* B108, 84-95(1931).—The diastase of the crust. style of the mollusc, *Pecten maximus*, exhibited no variation in its optimum pH with variation in the time of the expt. A decrease in the optimum temp. accompanied a decrease in the pH of the medium. When time and temp. were made equal to those in natural conditions, then the optimum pH was found to be the actual value in the living animal. JOSEPH S. HEPBURN

Proteolytic enzymes. VII. Peptidases of green malt. MASAKAZU SATO. *Compt. rend. trav. lab. Carlsberg* 19, No. 1, 1-26(1931), *c. C. A.* 23, 547.—The velocity of hydrolysis of the peptides alanylglycine (AG) and leucylglycine (LG) by the peptidases of malt indicate the presence of 2 different enzymes, and not one, as might be inferred from the data of Grassman and Klenk (*C. A.* 24, 863). LG, for which an enzyme prepn. has slight affinity, inhibits markedly the splitting of AG, for which the prepn. has great affinity. Calcn. according to the data of G. and K. would indicate that the LG affinity is about 32 times greater than the AG affinity, a conclusion which cannot be correct. W. D. LANGLEY

Asymmetric ester hydrolysis by enzymes. VI. The influence of cleavage products

upon the optical selectivity of an esterase. EUGEN BAMANN AND PAUL LAEVERENZ. *Ber* 64B, 897-909(1931), cf *C. A.* 25, 527-S—Recent objections (Rona, *et al.*, *C. A.* 25, 976) to the theory of esterase hydrolysis, based on the mass action law, have led to a review of the work. With liver esterase in *dl* mandelic acid ester mixts, the *d*-component was the more rapidly hydrolyzed, while in mixts. with the optically active components used separately as substrates, the *l*-component was favored. This action was explained as a result of the ratio of the velocities of the destruction of the intermediate products and the ratio of the enzyme affinities. B and L. also investigated the effect of alc. upon the optical selectivity of the esterase. They found, as expected, that the action of the *d*-ester was favored. 0.0575 g. 1.0N in 200 cc. reaction liquid decreased by 19° the negative rotation, of the mandelic acid produced. MARY E. LEAR.

Possible relationship between hemoglobin and chlorophyll as shown by the use of liver ext. ORAN RABER. *Science* 73, 457-8(1931)—Liver ext. was added to solns. of corn plants which were then kept in the dark 10 days. These were greener than the control plants. Some factor seemed to check the destruction of the chlorophyll. Whether this is the same effect as that observed in the use of liver ext. in cases of hemoglobin deficiency is the question. MARY E. LEAR.

The ultra-violet rays of artificial light. H. BUTZ AND TH. BOTTGER. *Z. Tierzucht. Zuchtungsbiol. Tierernähr.* 21, 48-72(1931)—The question of the usefulness of the rays in the prevention and cure of rickets is reviewed with special reference to animal husbandry practice. Data are presented showing that light treatment is without effect upon the hemoglobin content of the blood of rabbits. L. A. MAYNARD.

Enzymic synthesis of higher dextrans. S. NISHIMURA. *Biochem. Z.* 225, 261-6(1930)—It is shown that yeast autolyzate contains an enzyme which synthesizes amyllum. If 20 cc. of such a purified enzyme soln. is added to 100 cc. of 1% acbrodextrin soln. and the pH adjusted to 6.2 with 0.1 N HCl the mixt. gives a red to reddish violet color with I_2 after several days. At that time more enzyme prepn. is added and a few days later the soln. gives a violet to violet blue color with I_2 . Dextrin was also prepd. from the mixt., thus proving its synthesis from the lower mols. S. MOROULIS.

Influence of citrates on enzyme hydrolysis of starch. E. OMASOV. *Arch. intern. pharmacodynamie* 37, 98-107(1930), *Physiol. Abstracts* 15, 276-7—Citrates have no influence on diastases in optimum H ion concn., but they retard their action in high or low concn. of H ions. In the presence of chlorides citrates do not affect the action of pancreatic diastase, but in the absence of chlorides they increase this action when the H ion concn. is optimum. Citric acid does not hinder pancreatic digestion in the normal intestine. J. D. S.

Action of the alkaloids of cinchona on amylolytic enzymes. E. OMASOV. *Arch. intern. pharmacodynamie* 37, 108-14(1930); *Physiol. Abstracts* 15, 276—Quinine has no effect on pancreatic amylase. Quinine, cinchonine, cinchonidine and optochine retard the hydrolysis of starch by enzymes, quinine being the most strongly inhibitory. For taka-diastase the order is quinine, cinchonine, cinchonidine, optochine, while for pancreatic enzyme it is quinine, cinchonidine, optochine and cinchonine. J. D. S.

Oxidation mechanisms in animal tissues. M. DIXON. *Biol. Rev.* 4, 352(1929), *Physiol. Abstracts* 15, 501—This review is divided into the following sections: (a) the dehydrases, (b) peroxide formation, peroxidases and coupled oxidations, (c) Warburg's theory and the "respiratory enzyme", (d) cytochrome; (e) glutathione. A full discussion of the present position of our knowledge of these systems is given, together with extensive bibliography. J. D. S.

Viscosity of protoplasm as determining the rate of biological reactions. W. STILES. *Biol. Rev.* 5, 171(1930), *Physiol. Abstracts* 15, 495-6—This contribution to the problem of temp. coeffs. brings support to the views of Crozier and attacks those of Belchradek. The general applicability of the Sutherland-Einstein-Smoluchovski formula is contested by Stiles, who adduces several investigations on rate of diffusion in colloidal solns. of varying viscosity which show that increasing viscosity reduces the coeff. of diffusion relatively little. It is concluded that the viscosity of protoplasm, far from detg. the rate of biol. reactions, generally plays a very small, perhaps a negligible, part. J. D. S.

The distribution of urea in trichloroacetic acid filtrates as a function of the urea concentration of the medium. E. DELCOURT BERNARD. *Compt. rend. soc. biol.* 104, 784-5(1930), *Physiol. Abstracts* 15, 590, cf *C. A.* 25, 321—The distribution of urea in different portions of a CCl_3CO_2H filtrate from an albuminous fluid, blood, soln. of egg white, etc., is irregular. The differences are greater for blood than for albuminous fluids. The 1st parts of the filtrate from sheep blood contain the mean amt., then the curve rises, again falls and finally reaches the mean again. That is for a small concn. of urea.

When the conen is greater, the 1st portions have a low titer and the 2nd high.

J D S

Molecular weight of casein. III. D C CARPENTER *J Am Chem Soc* 53, 1812-26(1931), cf C A. 24, 1400—The protein contained in crude casein, which has been shown to have a mol wt between 75,000 and 100,000 by the ultracentrifuge, has been subjected to analysis. The S, P, cystine, tryptophan, tyrosine and histidine contents have been found to be 0.785, 0.856, 0.488, 1.237, 5.55 and 1.776%, resp. The most probable value for the mol wt of this protein, consistent with the above analyses and the results obtained with the ultracentrifuge, is 98,000. Details are given of the analytical methods used in detg the above components.

C. J WEST

Polarimetric reducing-sugar relationships of starch hydrolytic products resulting from diastatic action. D T ENGLIS, G T PFEIFER AND J L GABBY *J Am Chem Soc* 53, 1883-9(1931)—Theoretical consideration of the properties of the final products of hydrolysis of starch resulting from diastatic conversion seemed to indicate that the assumed const relationship between the polarimetric and reducing sugar values would not hold if glucose as well as maltose resulted from the action of the amylase system. It was further believed that the ratios found might furnish information on the mechanism of the hydrolysis of starch. Detns of the ratio have been carried out with taka-diastase as a glucose-forming amylase system. The ratios found were different from those anticipated and indicate that the changes in the early stages of hydrolysis are probably unlike those of malt diastase. While the ratios found were not const, detns of diastatic activity based upon the assumption of a const value will probably not involve an error more serious than those from other sources in such procedures. Improvements in the procedure will be necessary to give definite information as to the course of hydrolysis of starch.

C. J. WEST

Certain observations on the Donnan equilibrium (RONCATO) 2. The physical state of water bound by organic colloids and by the tissues (MARINESCO) 2. Properties of protein-cellulose membranes (VELLIZ, LOISELEUR) 2. Recent investigations on the constitution of pectins (BRIDEL) 10. Hydrolysis of gelatin (DESAI, BOLAM) 2. Röntgen diagram of collagen (HERZOG, JANCKE) 2. Physicochemical investigation of amino acids (TAKAHASHI, *et al*) 10. Further studies on the behavior of polypeptides containing proline toward the erepsin and trypsin kinase complex (ABDERHALDEN, ZUMSTEIN) 10. The behavior of polypeptides containing lysine with substitution in α - and ϵ -position, toward N alkali, erepsin and trypsin (ABDERHALDEN, SCHWEITZER) 10.

Collected Papers of The Mayo Clinic, 1931. Philadelphia W B Saunders Co 1130 pp \$13, net

KRAFFT, CARL F. Can Science Explain Life? Lancaster, Pa The Science Press Printing Co 94 pp \$1

WALDSCHMIDT-LEITZ, ERNST Vorträge aus dem Gebiete der Eiweisschemie. Leipzig Akad. Verlag 74 pp M 6 80

WEBSTER, RALPH W Legal Medicine and Toxicology. Philadelphia W B Saunders Co 862 pp \$8 50 Reviewed in *Military Surgeon* 68, 718(1931)

B—METHODS AND APPARATUS

STANLEY R BENEDICT

A new method for staining mycelium. A FERRARI *Boll soc intern. microbiol Ses stat.* 3, 26-7(1931)(in French)—The microscopic sections contg mycelium are kept for 20 min in the ruthenium red reagent (0.01 g of Ru NH₄ oxychloride in 15 cc H₂O). The mycelium color is red and the cell color is pale red, but if the sections are then transferred to a 10-20% KOH, the cell color disappears, while the mycelium color becomes red-yellow.

G A. BRAVO

A simple procedure for the determination of the hydrogen-ion concentration of normal human skin surface. B LUSTIG AND A PERUTZ *Arch Dermatol u Syphilis* 162, 129-34(1930).—A thin sheet of absorbent paper, moistened with distd water, is pressed for 5 min against the skin surface to be measured. The liquid is then expressed and its p_H is measured. To avoid error due to constituents of sweat, the p_H of the skin surface is detd after washing with water, then after washing with soap and finally after washing with C₂H₅. The detns of the p_H of the soln expressed from the absorbent paper were made with Wulff's indicator papers (These consist of a colloidal membrane, on which the indicator is adsorbed, $1/8-1/4$ of the paper is dipped in the soln to be measured, then wiped dry with filter paper, and compared with a scale.) The chief cause for a marked acid reaction of the skin is sweat. After treatment with water, soap

and CaH_2 , the p_{H} values found for various regions of the skin ranged from 5.1 to 5.8.

B. C. BRUNSTETTER

Modifications in technic for the determination of cholesterol, phosphorus and calcium on the same sample of blood serum. G. H. BARTL AND J. LABARER. *Trans. Roy. Soc. Can.* 24, Sect. V, 185-7 (1930).—Add to 1 cc. serum 1 cc. water and 1 cc. 2% $\text{K}_2\text{C}_2\text{O}_4$ mix, and after 20 min. centrifuge. Wash the ppt. with 2 cc. water, again centrifuge, dissolve the ppt. in dil. H_2SO_4 and titrate against 0.005 N KMnO_4 for Ca. Pour the first filtrate into a tall, somewhat conical 30-cc. beaker and add 5 cc. of 0.05% KOH in 60% alc., and 10 cc. Fe_2O_3 , shake vigorously and then allow to rest some minutes. Transfer the aq. layer to a test tube calibrated at 25 cc. Pour the wash water from the oxalate slowly into the ethereal layer, and after again the ethereal layer is added to the calibrated tube. Est. cholesterol in the ethereal soln. by Origan's method (C. A. 4, 2513, 2514), and the P in the aq. soln. by that of Benedict and Thies (C. A. 18, 3398). Comparison with the usual procedures showed accurate results. A. T. CAMERON

Determination of glycogen in liver tissue. ALBERT CARRUTHERS. *Chinese J. Physiol.* 5, 85-92 (1931).—A method is described for the estn. of glycogen in liver tissue which does not require the use of strong KOH , as does Pflüger's method. About 5 g. of chopped liver tissue is boiled for 1 min. in 10 cc. H_2O and then ground in a mortar. The mixt. is again boiled for 1 min. and made up to 50 cc. After centrifugation 20 cc. of supernatant fluid is pptd. with 5 cc. of 10% $\text{CCl}_3\text{CO}_2\text{H}$ and filtered. To an aliquot portion of filtrate EtOH is added to make a concn. of 70%. On the next day the ppt. is washed by centrifuging with 70, 95% and absolute EtOH and finally with ether. The ppt. is dissolved in boiling H_2O , HCl is added to make 2% and the mixt. hydrolyzed for 3 hrs. on the boiling water bath. After neutralization with NaOH sugar is detd. by a suitable micromethod. The values obtained by this method are somewhat higher than those by Pflüger's method, because of the presence of a carbohydrate which is insol. in 70% EtOH , but which is destroyed in 30% KOH . C. F. CORI

Determination of total sugar in liver tissue. ALBERT CARRUTHERS. *Chinese J. Physiol.* 5, 93-102 (1931).—The $\text{CCl}_3\text{CO}_2\text{H}$ filtrates described in the preceding abstract were used for the detn. of total sugar. An aliquot part was hydrolyzed for 3 hrs. in 2% HCl . Non-sugar reducing substances were removed from the neutralized soln. by pptn. with HgSO_4 according to West, Scharles and Peterson. No evidence was found for the presence of hydrolyzable carbohydrate other than glycogen. C. F. CORI

Folin's method for determining amino acids in the blood. P. M. RE AND D. PORICK. *Compt. rend. soc. biol.* 103, 1283-5 (1930).—See C. A. 25, 1856. E. J. C.

A simple method for the determination of assimilable pentosans. R. D. REGE. *Proc. 15th Indian Sci. Cong.* 1928, 51.—Treat a definite quantity of finely powd. material in the cold with 5% pure NaOH for 24 hrs. with occasional shaking. After neutralization distil the ext. with HCl by the usual Kroeber and Tollens' method. This method gives quite satisfactory results agreeing within 3% with those from the lengthy method (C. A. 21, 4062) previously described. E. J. C.

A micromethod for the analysis of proteins. N. NARAYANA AND M. SREENIVASAYYA. *Proc. 15th Indian Sci. Cong.* 1928, 163.—Van Slyke's well known method of protein analysis has been extended and modified for dealing with 100-200 mg. of substance, 0.05 of the quantity usually employed for a microanalysis. The pptn. of the hexone bases by phosphotungstic acid, which is affected by the presence of salts, is controlled by the effective removal of the HCl from the protein hydrolyzate *in vacuo* over freshly burnt lime. Arginine is estd. by flavionic acid, and histidine by bromination. Total monoamino and diamino is estd. by the Van Slyke micro app. Cystine is estd. by a microestn. of S according to Pregl. Tyrosine is detd. colorimetrically. E. J. C.

A simple method of protein determination in spinal fluid (and other body fluids). IVÁN BERGER. *Klin. Wochschr.* 9, 888-90 (1930).—The turbidity produced by sulfosalicylic acid added to the fluid is compared with that produced in known standards (casein solns.) in a specially designed comparator. The av. protein content of spinal fluid is 0.200%; the minimal value, 0.05-0.08%. H. EAGLE

Studies on blood volume. I. A micromethod for the indirect determination of absolute blood volume. ISTVÁN WENT. *Magyar Orvosi Arch.* 31, 39-53 (1930); cf. C. A. 23, 3942.—A colorimetric micromethod is described. A procedure is given for the elimination of errors caused by unequal distribution of dye. S. S. DE FINÁLY

Solution for the simultaneous staining of fat and nuclei. LAJOS KAROLINY. *Magyar Orvosi Arch.* 31, 440-1 (1930).—One hundred cc. of hematoxylin of Ehrlich is added to 20 cc. of acetone alc., then 2 g. of Sudan III is dissolved. The mixt. is shaken, let stand for 24 hrs., decanted and kept hermetically closed. Staining with this soln. is effected in 5 min. S. S. DE FINÁLY

Determination of calcium in blood serum. A. GRIGAUT AND I. ORNSTEIN. *Compt. rend. soc. biol.* 104, 747-9 (1930); *Physiol. Abstracts* 15, 573.—The serum is first evapd. to dryness in a Pt crucible and then incinerated to a white ash. This is washed with HCl and centrifuged, treated with NaOAc and $(\text{NH}_4)_2\text{C}_2\text{O}_4$, and again washed and centrifuged. After addn. of H_2SO_4 , titration with KMnO_4 is carried out to the point of persistence of a pink color. The nm. of cc. of 0.1 N KMnO_4 used multiplied by 100 gives Ca in mg per l. The amt. in normal serum is 95 to 110 mg by this method. G. G.

Microestimation of calcium. L. VELLUZ AND R. DESCHASEAUX. *Compt. rend. soc. biol.* 104, 976-7 (1930); *Physiol. Abstracts* 15, 563.—In estg. Ca, it is recommended to use org. solvents (alc., ether) to wash the Ca oxalate, which is sol. in distd. water. These solvents permit the removal of NH_4 oxalate and prevent mech. removal of part of the ppt. It is well also to oxidize the Ca oxalate in the cold by means of 150 N permanganate, of which the excess is estd. iodometrically. G. G.

Glass electrodes. IYAKUCHI YAGATA AND NOROSU FURUTANI. *Japan J. Gastroenterology* 3, 71-5 (1931).—The glass electrode can be employed in the detn. of the pH of biological fluids. C. M. McCAY.

The chemistry of uroselectan. ARTHUR BINZ. *Brit. J. Urol.* 2, 348-51 (1930); *J. Urol.* 25, 297-309 (1931), & *C. A.* 25, 2192.—Uroselectan, the Na salt of 5-iodo-2-pyridoneacetic acid, appears to be a satisfactory substance for use in visualization of the urinary tract. E. R. MAIN.

Methodical contributions. XL. A method for the determination of silver, ionized or bound to protein and of silver in organic substance. LUDWIG PRINGSSEN AND WADIM ROMAN. *Biochem. Z.* 225, 447-51 (1930), & *C. A.* 24, 3524, 4801.—The method consists in sepg. the Ag salts, Ag protein compd. and the metallic Ag and detg. the Ag in each fraction. Tissues are frozen with CO_2 and weighed. The frozen tissue is rubbed up fine at first dry, then with 5 cc. 0.1 N $\text{Na}_2\text{S}_2\text{O}_3$. The material is washed into a large centrifuge tube with 5 times 5 cc. 0.1 N $\text{Na}_2\text{S}_2\text{O}_3$. With blood, 5 cc. is treated directly with 30 cc. 0.1 N $\text{Na}_2\text{S}_2\text{O}_3$ in the centrifuge tube. This is now kept overnight in a cool, dark place, and sharply centrifuged. The turbid soln. is decanted into a second centrifuge tube, placed in a water bath at 70° to coagulate the proteins and again centrifuged. The clear soln. is now transferred to an Erlenmeyer flask, and the residue once more treated with 5 cc. 0.1 N $\text{Na}_2\text{S}_2\text{O}_3$. The combined $\text{Na}_2\text{S}_2\text{O}_3$ exts. contg. the inorg. Ag compds. constitute fraction A. The residue is mixed with 5 cc. 5% NH_4OH and after standing overnight is also centrifuged for about 20 min. and treated 4 times more with 5 cc. of the NH_4OH soln. The combined centrifuged NH_4OH exts. constitute the B fraction contg. the Ag bound to the protein. The last residue, fraction C, contains the metallic Ag. It is evapd. to dryness on the water bath, it is then charred on the sand bath. This is also done with the fractions A and B. When cool the charred material is taken up with halogen-free HNO_3 and perbydrol and evapd. to dryness; this process is repeated until a colorless soln. is obtained. All org. material is thereby oxidized and Ag converted to AgNO_3 , while the halogens are driven off. The white residue is now dissolved in H_2O and evapd., this treatment continued until the HNO_3 is completely removed. The residue is then acidified with a few drops HNO_3 (1:3), and titrated with 0.001 N NH_4SCN in the presence of a crystal of iron alum, each cc. corresponding to 107.9 g. Ag. S. MORGULIS.

Manometric determination of catalase. AKIJI FUJITA AND TAKESHI KODAMA. *Biochem. Z.* 232, 20-34 (1931).—The Warburg manometric app. is used to det. the catalase activity of bacteria. The catalase coeff. = Q_{cat} = (cc. O_2 produced in 30 min. at 38°) \div (mg. dry wt.). If the rise in level of manometer is h mm. and the vessel const. k_{cat} sq. mm., the O_2 formed x_{cat} is calcd. from the formula $x_{\text{cat}} = h k_{\text{cat}}$. The Q_{cat} for *Staphylococcus pyogenes albus* was found to be 5120-5200; for gonococci 8700-8720; for *Saccharomyces saké* 44.5-45.2; for rabbit retina 65-53; for Jensen's rat sarcoma 119-113; for rat lung emulsion 475-459; and for a purified catalase prepn. 900,000-895,000. The optimum activity was found to be between pH 6.4 and 8.0. NaCl did not produce an inhibitory action nor was any depressing effect observed with a phosphate concn. below 0.005 mol. No difference in the catalase activity has been found in enterococci cultured aerobically or anaerobically. On the contrary in the case of bacteria no catalase was found in the anaerobes, although some anaerobes like the Shiga dysentery bacillus are also free from catalase, and facultative anaerobes contain a moderate amt. of catalase. Two long lists are reported of the Q_{cat} values for various bacteria and tissues from several organisms. S. MORGULIS.

Application of the Folin method for determining amino nitrogen in 0.2 cc. of blood. S. A. POVOINSKAYA. *Biochem. Z.* 232, 69-71 (1931).—The reagents and procedure are the same as in the original Folin method except that the blood (0.2 cc.) is obtained from

the finger by means of a micropipet, and only 0.2 of the amts of the reagents are employed throughout the detn.

S. MORGULIS

A simple colorimetric pH determination. E. HERZFELD *Biochem Z* 232, 72-7 (1931).—Buffer mixts of Na_2HPO_4 , $2H_2O$ and citric acid with a range from pH 2.2 to 8.0, and mixts of H_3BO_3 , KCl and $NaOH$ with a range from pH 7.8 to 10 are employed together with 7 sep indicator solns. The detns are carried out in the 1931 model of the Leitz Universal colorimeter in the following way. After adding the specified quantity of indicator to 100 cc of buffer mixt this is divided into 3 parts. One is placed in the lower left cup. The second portion is acidified with a few drops of concd HCl and is used in the right upper compensating cup. The third portion made alk with a few drops of $NaOH$ is placed in the lower right cup, while the left compensating cup is filled with distd H_2O . The depth of the buffer mixt is set at 10 mm and the depths of the cups with the acidified and alkalinized mixts are so adjusted that a perfect color match is obtained. The relative depths of the last two cups corresponding to each pH between 2.2 and 10.0 are tabulated. The pH values of urine, serum, etc., are detd in a similar way by adding the proportional quantity of indicator and proceeding as before. With the microcolorimeter the detns can be made with only 10 cc of the original sample, and from the depths of the acidified and alkalinized portions which together match the color of the untreated sample the pH is read off from the table. The color of the material studied does not interfere in the detns by this method.

S. MORGULIS

Titrimetric determination of fat and lipid substances by the Bang chromic acid oxidation method. H. STRAUM *Biochem Z* 232, 128-45 (1931).—The reduction quotient of Bang's e , the cc. 0.1 N $K_2Cr_2O_7$ necessary for the oxidation of 1 mg fat or lipid material has been redetd on a series of substances, and has been found to be generally lower than that given by Bang. This quotient was found to be larger the greater the excess of $K_2Cr_2O_7$ employed. On the other hand if the quantity of fat exceeds about 0.4 g the value of the quotient tends to decrease. The quotient is also affected by the temp of the reaction mixt. after the addn of H_2SO_4 by the length of time during which the $K_2Cr_2O_7$ - H_2SO_4 mixt has been kept, etc. For triolein, tripalmitin, tristearin, the quotient is 1.22, for oleic and palmitic acids 1.87, for stearic acid, 2.03, for cholesterol, 2.37, for cholesterol palmitate, 2.16, for cholesterol stearate, 2.27; and for lecithin, 1.44. The method has been applied to quantities ranging from 0.5 to 10 mg. The lipids in a 100 cc Erlenmeyer flask is treated with 30 cc. N $K_2Cr_2O_7$, then with 15 cc concd H_2SO_4 , and the mixt is left standing until it cools to room temp. This is now transferred with 300 cc H_2O to a 500 cc flask and after the addn of 10 cc 10% KI is titrated with 0.1 N $Na_2S_2O_3$. The results of such analyses with the different pure substances are tabulated. The results are also plotted and the amt of substance corresponding to the cc 0.1 N $Na_2S_2O_3$ used up can be read off directly from the curves with an error of 2.92%.

S. MORGULIS

Dialysis procedure for the determination of serum proteins. A. A. SCHWARTZ AND KLARA TULICHINSKA *Biochem Z* 232, 323-34 (1931).—One cc. of serum is dild with 2 cc. H_2O and 3 cc satd $(NH_4)_2SO_4$, thoroughly mixed and left standing 1/2-1 hr. The pptd globulins are collected on a small filter paper and washed 3 times with 2 cc. of 1/2 satd $(NH_4)_2SO_4$ soln, the wash water being added to the original filtrate. The globulins are transferred to a small dialyzing shell by dissolving them in small quantities of H_2O and allowing the soln to drip through the funnel, and the filter paper is washed with 5 cc of 250 N $NaOH$, 1 cc. is used at a time. The $NaOH$ is then neutralized by adding from a buret an equiv quantity of 0.1 N HCl , the shell is then closed with a stopper and the dialysis carried out. The filtrate contg the albumins is also put into a dialyzing shell. The dialysis is carried out for about 8 hrs at a temp of 45-60°, against water circulating at the rate of 10 l per hr until the reaction with Nessler's reagent is neg when distd water is passed at the rate of 1/2 l per hr. In the shell with the globulin soln at the end of dialysis the cuglobulin is present as a firm ppt while the pseudoglobulin is in soln and the 2 fractions are quantitatively sep'd by centrifuging. The cuglobulin is washed with 3-5 cc H_2O which is placed together with the original pseudoglobulin soln in a 25 cc flask. The ppt is dissolved in 5 cc. $N/250$ $NaOH$ and also transferred to a 25 cc flask. The dialyzed albumin soln is entirely clear and is put into a 50 cc flask, to which is added also the $N/250$ $NaOH$ used for rinsing the shell. The N is detd on these 3 portions and expressed in terms of mg per cc serum.

S. M.

A method for determining lactose in the urine. E. FREUND AND B. LUSTIG *Biochem Z* 232, 449-51 (1931).—To 100 cc urine add 25 cc HNO_3 and evap to 20 cc. Dil and cool the mixt and wash the pptd mucic acid with alc. and ether, dry and weigh. Considerable amts of mucic acid are found after feeding 40-100 g lactose or galactose, but none is obtained when 1-2 g lactose is added directly to the urine, which suggests

that the mucic acid is derived from the galactose. However, a 1% aq soln of lactose also yields much mucic acid. The reason for the neg results with urine is found in the fact that urea and NH_4 salts prevent the formation of mucic acid under the same exptl conditions. A modified method was, therefore, developed for the lactose detn. In urine add 1-g lactose to 100 cc urine, acidified with 0.1% AcOH , and treat with 10 cc of 10% Pb acetate soln. Wash the ppt with 0.1% AcOH , and remove the excess Pb in the filtrate by adding, drop by drop, aq H_2SO_4 . Neutralize the filtered soln with NaOH , make slightly acid with AcOH and carefully evap. to a small vol. Boil the residue 2-3 times with 100 cc alc., cool and filter. This procedure removes urea, NH_4 salts and sugar except the lactose which is practically insol in cold alc. Dissolve the alc-extr. residue in H_2O and filter. Lactose detns made by this procedure, the lactose being detd polarimetrically or by reduction methods, give recoveries with 10% of the added amt.

S. MORGULIS

Determination of Iron in organs and in body fluids. HANS HORSTERS. *Biochem Z.* 232, 469-78(1931).—Organs are oxidized by the Neumann method, urea being added to destroy nitroso compds. The ash soln is added with 1-cm water to 50 cc in case 2-5 g of tissue was oxidized. To this is added for every 2 mg Fe 10 cc of Neumann's Zn reagent, which is made by dissolving separately 25 g ZnSO_4 and 100 g Na_2HPO_4 , mixing the two in a 1-l flask, redissolving the ppt with purest aq H_2SO_4 and dilg the soln to 1 l. The ppt. double salt of Fe and Zn phosphate is carefully dissolved in pure NH_4OH (d 0.90) added drop by drop, and the clear soln is heated 30-60 min. The ppt. is filtered off hot and washed several times with hot water. The ppt. is now dissolved on the filter paper in exactly 5 cc purest HCl (d 1.124), and the paper is washed several times with hot water, but the total vol should not exceed 20 cc. The soln of the double salt, which under certain conditions may be purified by a second pptn with the Zn reagent, is transferred to a 50-cc volumetric flask and filled to the mark with 30 cc of 40% NH_4SCN . The colorimetric detn is then carried out after 3 min. as standard, a suitable soln of FeCl₃ (0.10 to 0.01 mg % Fe) acidified with HCl is used. The standard soln should not be used longer than 20 min, when it should be replaced by a freshly prepd soln.

S. MORGULIS

Determination of methylglyoxal, pyruvic acid and acetaldehyde separately or in mixture as well as a determination of lactic acid in the presence of these substances and in a single sample of material. FRANK SIMON AND CARL NEUBERG. *Biochem Z.* 232, 479-81(1931).—Hd fluids contg methylglyoxal, AcH and pyruvic acid are deproteinized in an acid medium, e.g., $\text{CCl}_3\text{CO}_2\text{H}$. These are pptd with 2,4-dinitrophenylhydrazine, the hydrazones being either filtered off or centrifuged off. For lactic acid detns the excess 2,4-dinitrophenylhydrazine can be removed first by making the soln alk and filtering off the ppt. The traces of reagent still remaining in soln can be extd with LiOAc . However, where carbohydrates must be removed by the Cu-Ca pptn method the extn is unnecessary because the traces of the 2,4-dinitrophenylhydrazine are carried down in the ppt. In this filtrate lactic is detd by the usual procedures. The pptn of methylglyoxal, pyruvic acid and AcH is practically quant. for solns contg 0.1 to 1.0%, and this is also true when all 3 substances are present together, but in the presence of AcH the pptn must be made in the cold. The suspended ppt. is then left for 10 hrs at 37° in a glass-stoppered Erlenmeyer flask. After cooling this is filtered through a Goret glass filter No. 1G-1ab, washed 3 times with cold 2 N HCl and then with cold water. For the sepn of the dinitrophenylhydrazones advantage is taken of the fact that only the pyruvic acid hydrazone is sol in Na_2CO_3 and can be dissolved out from the ppt. in the glass filter. Since this soln is colored brown it is possible to tell when the extn is completed because the soln begins to come through colorless. The hydrazone is again pptd in the filtrate by carefully acidifying with HCl , filtered off, washed, dried at 110° and weighed. The sepn of the other 2 hydrazones is effected by means of hot 91% alc in which the AcH deriv alone is sol. The extn, is, therefore, again carried out in the original glass filter by passing hot alc until the soln which is at first yellow becomes colorless. The alc soln of the AcH hydrazone is evaporated in a weighed beaker, the residue is dried in a vacuum desiccator to const wt and weighed. The methylglyoxal hydrazone is washed on the filter with ether, dried at 140° and weighed. The method is also applicable to the detn of traces since these yield methylglyoxal on distn with 20% H_2SO_4 . By treating the methylglyoxal deriv with alc. or Na_2CO_3 even levulinic acid and furfural coming from hexoses and pentoses, resp., do not interfere.

S. MORGULIS

Calcium and constitution. New method for the estimation of calcium in the blood. LUIGI F. LUNARDONI AND GIUSEPPE MIRAGLIA. *Rass clin terap sci affini* 30, 1-7 (1931).—Two cc. of serum is mixed in a centrifuge tube with 2 cc of water and 1 cc. of

said $(\text{NH}_4)_2\text{C}_2\text{O}_4$ soln. The mixt. is allowed to stand for 4 hrs., centrifugalized for 1 min., decanted and the ppt. is washed with 4 cc. of 1% $(\text{NH}_4)_2\text{C}_2\text{O}_4$ soln. The mixt. is centrifugalized, decanted and the ppt. is dissolved in the smallest possible quantity of 4% HCl. One drop of alc. phenolphthalein soln. is then added, enough of 2% NH_4 to produce a red color, and 1 cc. of said $(\text{NH}_4)_2\text{C}_2\text{O}_4$ soln. After 2 hrs. the ppt. is conveyed to an ashless filter, 2-3 washings of 3 cc. of 1% $(\text{NH}_4)_2\text{C}_2\text{O}_4$ soln. are employed to bring the last traces of the ppt. on the filter. The filter is placed in an oven at 100° , until the ppt. is almost dry. Then the filter with the ppt. is conveyed to a Pt crucible and incinerated. The residue of CaO is dissolved in 5 cc. of 0.01 N HCl, the soln. with washings is conveyed to a beaker and the excess of HCl is titrated with 0.01 N NaOH with methyl red as an indicator. A blank test should be run to det. the alk. of the ashes of the filter paper used. $\text{Calcn. } Q = 50(5 - a - y) \times 0.0002$, where $Q = \text{Ca in mg. } \%$, $y = \text{no. of cc. of 0.01 N HCl used for the neutralization of the ashes originated from the ignition of the filter and } a = \text{no. of cc. of 0.01 N NaOH used for the neutralization of HCl}$. G. SCHWACH

Determination of uric acid in blood. F. MONTREUIL *Annales soc. chim. su. guim.* 29, 264-6(1931).—The method depends on the reduction of $\text{K}_2\text{Fe}(\text{CN})_6$ by uric acid in the presence of Na_2CO_3 , the $\text{K}_2\text{Fe}(\text{CN})_6$ formed giving a blue color with FeCl_3 . Blood is dealbuminized by the Folin method, by adding Na tungstate and 0.66 N H_2SO_4 . Uric acid is isolated as Ag urate, uric acid is set free and oxidized in an alk. $\text{K}_2\text{Fe}(\text{CN})_6$ soln., the soln. is acidified with HCl and FeCl_3 is added to give a blue color. Deto. is by color comparison. R. M. STONES

LANHAUD, H. *Le lipiodol en urologie*. Paris: N. Maloine, 71 pp. F. 15

Fixing solution for histological purposes. WALTER SCHILLER. *Ger.* 522,317, June 19, 1930. See *Austrian* 119,237 (*C. A.* 25, 721).

C-BACTERIOLOGY

CHARLES B. MORREY

Necessity of the control and standardization of methods of investigation of vitamin products. EDMUND LULESZ AND A. PRZEDRZECKA JEDRZEJOWSKA. *Med. Doświadczalna i Społeczna* 12, 357-87(1930).—The popularization of the science of vitamins is discussed. Control tests as performed by the Biochem. Department of the State Hygienic Institute of Poland are described. J. KUTNER

The kinetics of the bacterium-bacteriophage reaction. ALBERT P. KUTNER AND JOHN H. NORTHCROP. *J. Gen. Physiol.* 14, 223-54(1930).—Expts. were made on 16-hr. cultures of *Staphylococcus aureus* and antistaphylococcus bacteriophage in nutrient broth of pH 7.0 at 36° , the cultures being mechanically shaken. The following symbols are used in explaining the results: B = bacterium, P = bacteriophage, $[B]$ = bacteria/ml., $[P]$ = phage/ml., B_0 = initial value of $[B]$, B_e = $[B]$ at max. stationary phase of B growth, P_0 = initial value of $[P]$. B growth in P - B mixts. resembles growth in controls without P except where the initial P/B ratio is very high. 1. destruction of B does not begin soon after mixing P and B . B growth is not stimulated by P but it is an essential conditioning factor for P formation. B growth and P formation exhibit short lags during which time P diffuses so rapidly into B that by the end of the lag period 70-90% of the total P is associated with B , the rest being extracellular. During the logarithmic B growth phase, the formation of P is also logarithmic but at a higher rate. Therefore it is not true that each time a B divides a certain amt. of P is formed. When B growth shows pos. acceleration equal between extracellular and intracellular P fractions becomes established, P is maintained until the onset of lysis. Extracellular $[P]$ constitutes a small const. % of total $[P]$. The distribution of P suggests that of a simple chem. compd. rather than that of a complex organized parasite. Lysis begins when $\log P/B = 2.1$. If this value is reached or exceeded when P and B are mixed B does not grow and lysis soon occurs. Lysis is probably brought about by a particular $[P]$ per B and not by a certain $[P]$ per ml., but the major conditioning factor is in doubt. The max. $[P]$'s of lysates made by mixing a const. initial $[B]$ with widely varying P_0 's fall within a narrow range. The lytic destruction of B is logarithmic with time, being similar in this respect to most death processes. Cells need not be growing to undergo lysis. During B lysis a considerable percentage of total max. P formed is destroyed, the loss is chiefly in intracellular P . From 70 to 90% of the final P present at the completion of bacteriophagy is set free during bacterial soln. Although B growth is essential for P formation, the different temps. of B and P formation suggest that the 2 processes are basically different. A

similar interpretation is given for B soln and P inactivation. The $[B]$ at which lysis occurs for given values of B_0 and P_0 may be calcd from the equation $\log B = \frac{2.1 + n \log B_0 - \log P_0}{n - 1}$. If this value of $\log B$ is substituted

in the equation $t = \frac{1}{k B_0} \left[\log \frac{B(B_0 - B)}{B_0(B - B_0)} \right]$ gives satisfactory agreement with

exptl values for lysis time. Kinetic analysis of the $P - B$ reaction predicts a straight line relation when $\log P_0$ is plotted against t lysis for a const B_0 . The plot has been used in a method for the estn of P (cf C A 24, 3810) on the basis of exptl observation. Its use is made more rational by the facts given in this paper.

C. H. RICHARDSON

Natural and chemically defined media. A. BERTHELOT *Bull soc chim biol* 12, 1025-8(1930).—Various media suitable for the growth of *M. tuberculosis* are discussed. A high content of arginine is favorable, and use may be made of products derived by proteolysis from naturally occurring substances.

B. C. A.

Occurrence of cellulose-decomposing bacteria of the group Cytophaga in the rumen of the ox. CH. ARMAND *Boll soc intern microbiol Ser ital* 3, 35-40(1931) (in French).—In the ox rumen, 2nd and 3rd stomach and in the cecum (examd by the Winogradsky method (C A 24, 3309)) several species of cellulose-decomposing bacteria were found, and *Cytophaga hutchinsonii* was isolated. The oxidation products of cellulose give the reaction of oxycellulose, but do not reduce Fehling's soln.

G. A. BRAVO

The relation between the acidity of milk serum and its bacterial content. M. BONDROIT *Boll soc intern microbiol Ser ital* 3, 77-8(1931) (in French).—The coccus groups of microorganisms, which are weak acid producers and very slightly resistant to acid, are predominant at first in milk serum, but afterward they are gradually replaced by organisms of the bacillus group, which produce more acid and are more acid resistant.

G. A. BRAVO

Favoring influence of local disturbances caused by adrenaline on the development of bacterial infections. M. RENAUD AND A. MIGET *Compt rend soc biol* 103, 1052-4(1930), *Physiol Abstracts* 15, 412.—Adrenaline favors the growth of microbes in living tissues. Subcutaneously injected before inoculation, it induces rapid development of pathogenic organisms, and saprophytic bacteria which are ordinarily quickly destroyed continue to live in the prepd. regions. Added to the culture medium, adrenaline also renders bacteria more active than those cultivated without it.

G. G.

Egg-yolk agar medium for the growth of tubercle bacilli. RUSSELL D. HERROLD *J Infectious Diseases* 48, 236-41(1931).—Approx. 15% raw and sterile egg yolk added to melted ordinary beef ext. agar at pH 7.5 forms an excellent medium for the isolation of tubercle bacilli from body fluids and tissues.

JULIAN H. LEWIS

Further studies on bacterial lipoids. M. GUNDEL AND W. WAGNER *Z Immunitäts* 69, 63-76(1930).—The acetone-sol. fraction of the alc. ext. of *Pseudomonas aeruginosa* was bactericidal for all of a series of bacteria except the typhoid colon group. The same fraction from other bacteria was bactericidal for anthrax bacilli alone. The bacteria are killed by bacteriolysis. The bactericidal properties of bacterial lipoids are thermostable in an aq. medium, but are inactivated by heat when in a protein medium. Bactericidal action is dependent on the concn and age of the lipoids. After standing and after exposure to light, the lipoids are decreased in their bactericidal properties but not in their antigenic action. Immunization with lipoids from *P. aeruginosa* produces ipoid antibodies, while lipoids from staphylococci and typhoid and colon bacilli are not antigenic. The bactericidal and antigenic properties of bacterial lipoids are not identical. Fatty acids and their esters and lipoids from plant but non-bacterial origins are bactericidal.

JULIAN H. LEWIS

The chemical differentiation of hemolytic and nonhemolytic streptococci. LOUIS AVILES *Z. Immunitäts* 69, 433-7(1931).—By means of the Berlin blue reaction the changes in the hemoglobin of blood agar plate cultures of streptococci were studied. In the clear zone around hemolytic streptococci hemoglobin is absent and the bacterial colonies give the reaction for Fe. In nonhemolytic cultures the zone around the colonies gives a reaction for Fe while the colonies themselves do not.

JULIAN H. LEWIS

Are there gas-forming variants of typhoid bacilli? CURT SONNENSCHNIG *Z. Immunitäts* 69, 449-63(1931).—A critical study of the claims of various workers to have observed variants of typhoid bacilli that form gas leads to the conclusion that these claims are unwarranted. These organisms never acquire this property by any process of regression, transformation or disson. Gas formation by an organism excludes its possibility of being *Eberthella typhi*.

JULIAN H. LEWIS

A quantitative study of the respiration of *Staphylococcus aureus* cultures lysed by bacteriophage. MARGOT D. FAYON, JR. *J. Bact.* 21, 147-50 (1931).—When a large amt. of bacteriophage actively lyses a culture of staphylococcus in the stage of max. rate of growth the observed rate of O consumption is greater than can be accounted for by the number of intact cells present. A lysed culture contg. negligible numbers of living bacteria continues to give off CO₂ and take up O for several hours. Either the bacteriophage itself or some product of the action of bacteriophage on or in combination with bacteria respires. Lysis by bacteriophage is not a simple autolysis. J. T. M.

Isolation and study of an apparently wide-spread cellulose-fermenting anaerobe, *Cl. cellulosolvens* (N. Sp.?). PHILIP B. COWLES AND LEO I. RITTIGER. *J. Bact.* 21, 167-82 (1931).—This organism was isolated from horse feces and is morphologically similar to the organisms described by Omeliansky. The rods are thin, often slightly curved, and form spherical terminal spores. Colony formation can best be obtained on a peptone beef infusion agar contg. dextrin and cysteine although it is not selective. The best liquid medium is beef infusion peptone phosphate agar contg. a strip of filter paper. Cellulose, dextrin, arabinose and xylose alone of the materials tested are attacked. Glucose is not utilized, a fact in contradiction to the so-called theory of carbohydrate gradients. Dextrin or glucose has little effect on cellulose decompo. The products of cellulose decompo. by this organism are CO₂, H₂ and org. acids.

Influence of osmotic pressure upon spore germination. HAROLD R. CLARRAN. *J. Bact.* 21, 197-209 (1931).—There is a definite limiting concn. of food below which no germination occurs, irrespective of the osmotic pressure. There is no evidence of a purely osmotic limitation in low pressure solns., limitation being probably due to food deficiency. Germination is more rapid and complete at relatively low pressures. When the min. nutritional requirements are exceeded, the increased concn. of food may retard germination because of an unfavorable increase in osmotic pressure. A pressure of 36 to 40 atm. becomes inhibitory.

Peptone water with rhamnose as a differential culture medium for *Pasteurella pestis* and *Corynebacterium rodentium* Pfeiffer. A. Bessonova. *Zentr. Bakt. Parasitenk.*, 1 Abt., 119, 33-5 (1930).—*C. pseudotuberculosis* produces acid, *P. pestis* does not.

Investigations on the power of the *Brucella melitensis* group of bacteria to produce hydrogen sulfide. The production of hydrogen sulfide as a criterion for the differentiation of the members of the *Brucella* group. GIOVANNI PAVILLI. *Zentr. Bakt. Parasitenk.*, 1 Abt., 120, 24-34 (1931).—See C. A. 24, 5329.

The influence of bile on the agglutinability of a typical peristaphyroid bacilli. N. G. KLUYEVA. *Zentr. Bakt. Parasitenk.*, 1 Abt., 120, 34-40 (1931).—Some inagglutinable strains become agglutinable when grown in a bile-contg. medium. Spontaneously agglutinable strains are not changed by bile.

The frequent occurrence of streptococci in human group which produce excessive polysaccharides when grown on the surface of a substrate containing sucrose or raffinose. J. ØRSKOV AND K. A. FOLSEN. *Zentr. Bakt. Parasitenk.*, 1 Abt., 120, 125-7 (1931).—This is a const. property of many strains.

A substance which increases bacteriophage action. FRIEDRICH HODER. *Zentr. Bakt. Parasitenk.*, 1 Abt., 120, 162-5 (1931).—There is a substance in human and animal blood which increases bacteriophage action. It appears also to be present in certain bacteria. Its nature is not understood. It may be enzyme like or vitamin-like in character.

Bacteriological study of chronic infectious arthritis. JOHN W. GRAY AND CECIL H. GOWEN. *J. Med. Soc. New Jersey* 28, 38-41 (1931).—G. and G. describe a modification of the medium used in Cecil's method (*Arch. Internal Med.* 43, 571) for the isolation of a streptococcus from the circulating blood of patients with chronic infectious arthritis. The beef heart infusion is prepd. as described by C. At the end of the boiling and filter mg. 0.5% NaCl, 1% gelatin, 1% glucose and 1.5% Witte peptone are added. The mixt. is kept in the Arnold app. for 20 min., titrated to pH 8, placed again in the Arnold app. for 60 min., and retitrated. The pH should not be below 7.8. The soln. is filtered through paper and sterilized in the Arnold app. for 30 min. on 3 successive days in bottles contg. a teaspoonful of powd. CaCO₃. If the pH is 7.6-7.8 the medium is satisfactory. With this improved medium it is possible to produce growth in the form of diffuse clouding of the medium in 1-4 days without opening the bottles. C. SCHWOCK.

Investigations on the agglutination of various species of *Leishmania*. ALESSANDRO LAURINSCHI. *Pediatrics Rinnia* 39, 345-50 (1931).—The expts. were carried out on 4 strains of *L. infantum*, 2 strains of *L. donovani*, 1 strain of *L. tropica* and 1 strain of *L.*

brasilensis Immune sera were prepd by injecting rabbits with cultures of these strains. The immune sera from the rabbits injected with *L. donovani* agglutinated the homologous strains and also the *L. infantum* strains. The immune sera from *L. infantum* agglutinated the homologous strains and also the strains of *L. donovani*. The immune sera prepd with *L. tropica* and *L. brasilensis* agglutinated only the resp homologous strains. *L.* assumes on the basis of his results that *L. donovani* and *L. infantum* are identical or at least very closely related. G. SCHWOCH

Specific and nonspecific polysaccharides of type IV pneumococcus. MICHAEL HEIDELBERGER and FORREST E. KENDALL. *J. Exptl. Med.* 53, 625-39 (1931), cf C. A. 24, 653. — Three N contg polysaccharides have been isolated from autolyzed cultures of Type IV pneumococcus: a type sp carbohydrate (I), differing markedly from those of Types I, II and III pneumococcus and representing a type of substance hitherto not observed among sp polysaccharides; a chemically similar carbohydrate (II) without sp function, and the "C" substance (III), or species-sp polysaccharide of Tillet, Goebel and Avery (C. A. 25, 511). Detailed directions are given for the sepn of these 3 compds. That I was not obtained pure is evidenced in the variation of $[\alpha]_D$ from 17.0° to 35.0°, in the acid equiv from 1250 to 3330, in the total N from 4.7 to 5.9, in the reducing sugars on hydrolysis (as glucose) from 67 to 76.4, 2 samples contained in % 0.1 and 0.4 amino N, 5.6-5.9 acetamido N, less than 0.1 P and 1 sample showed 45.9 C and 6.7 II, the hydrolysis products are AcOH and an amino sugar deriv. II showed $[\alpha]_D$ 10°, acid equiv 4540, total N 5.9 acetamido N 5.6, reducing sugars on hydrolysis (as glucose) 55 (av. values), the hydrolysis products are glucosamine, amino sugar deriv. and AcOH. III gave as av. values, $[\alpha]_D$ 42°, acid equiv 1050, total N 6.1, amino N, 0.9, acetamido N 3.7, reducing sugars 36, hydrolysis products, amino sugar deriv., H_3PO_4 and AcOH. III contained 4% P, at 100° H_3PO_4 is only slowly split off by N HCl or NaOH. The relationship of these compds. to chitin is pointed out and its bearing indicated on the unsettled controversy as to whether or not chitin occurs in bacteria. C. J. WEST

Metabolism of S and R forms of pneumococcus. PHILIP FINKLE. *J. Exptl. Med.* 53, 661-76 (1931). — Results are given of studies on the respiration and glycolytic metabolism of pneumococcus Types I (f), II (II) and III (III) and of the R forms derived from these. The metabolism of the S and R forms is compared and the relationship between changes in virulence, changes in chem. constitution and changes in metabolism is discussed. There is no respiration in Ringer soln unless sugar is added. The pH (7.8) that is optimal for growth of pneumococcus is also the pH at which the max. respiration occurs. The intensity of respiration varies with the type in the strains used. The respiratory capacity of I is 56% of that of III, which in turn is 71% of that of II. The anaerobic glycolysis is approx. the same for all 3 groups. I is capable of aerobic glycolysis, while II and III do not effect glycolysis aerobically. The energy set free in respiration is considerably greater than that set free in glycolysis. The oxidation quotient for lactic acid is of the same order as that found by Meyerhof in muscle and by Warburg for mammalian tissues. The respiratory capacities of I and III are changed on conversion of the smooth to the rough form. For I the respiration is increased 110%; for II it is diminished 45%, for III there is only a slight diminution in respiratory activity (16%). The anaerobic glycolysis is increased 25% on the av. for all R forms irrespective of type derivation. I on being converted to the R form loses its capacity for aerobic glycolysis, while III gains it. The O consumption by pneumococcus compared with that of the human tubercle bacillus and of mammalian tissue, for the same time intervals, wt for wt, is I consumes 13 times as much O as does the tubercle bacillus (H. 37), II consumes 34 times as much O, and III consumes over 20 times as much O as does isolated rat kidney tissue and almost 100 times as much O as does isolated dog muscle. C. J. WEST

Dyes [for coloring bacteria] (Fr. pat. 698,076) 25.

FRETS, G. P. Alcohol and the Other Germ Poisons. The Hague. Martinus Nijhoff. 179 pp. Fl. 6, bound, Fl. 7.50.

D—BOTANY

THOMAS G. PHILLIPS

The possible presence of norleucine in the proteins of castor-oil seeds. R. NUCCORINI (with G. BAGNIS). *Boll. ist. super. agrario Pisa* 6, 227-37 (1930). — Among the hydrolysis products (with coned boiling HCl) of the proteins of germinating castor-oil seeds, together with ricinine, a crystal compd. was found, which had the compn.

$C_{11}H_{19}O_4N$, m 270–80° (with partial decompn), $[\alpha]_D -16.52$ in aq soln. and +14.09 in 20% HCl soln. This compd was racemized and the Bx deriv (prepd by the Fischer method, cf *Ber* 33, 2370(1900)) had m p 134–5° (with previous softening). The yield was too low to continue the identification expts, but very likely the compd is norleucine, which till now has not been found among the hydrolysis products of the vegetable proteins.

G. A. BRAVO

Fluorescent substances contained in plants. II. R. NUCCORINI (WITH M. MONSACCHI) *Boll ist super agrario Pisa* 6, 252–62(1930) —The plants examd. (olive tree, black cabbage, rape, spinach, lettuce, celandine, rue, sweet clover) contain water-sol substances fluorescing bluish to Wood's light (3950–4750 Å U.), which are probably pigments formed as secondary products of the biochem activity of the plants. In several plants characteristic fluorescing compds are present, e.g., *oleinola* (fluorescing green blue, 4500 Å U.) in the olive tree and *berberine* (fluorescing gold yellow, 5500 Å U.) in the celandine. The fluorescent substances are not injured by the action of the molds.

G. A. BRAVO

Studies in sap analysis. 1. Carbohydrates in saps. M. SARENIVASAYA AND H. N. SASTRY *Proc 15th Indian Sci Cong* 1928, 163–4 —The estn of sugars in saps is complicated by the presence of interfering impurities such as tannins, gallic acid, dextrins, proteins and similar substances. Iptg agents like basic Pb acetate, dialyzed Fe, fibrous alumina, $HgCl_2$, etc., have proved unsatisfactory in one way or another. A very suitable method consists in the preliminary removal of dextrins and the greater part of the proteins by an addn of 10 vols of 95% alc contg a little NH_4 . The alc filtrate is concd on a water bath; tannins, etc., are eliminated by dialyzed Fe; and the clear filtrate is used for the sugar estns. by Bertrand's method. The direct reducing sugars are estd as dextrose. The citric acid hydrolysis gives sucrose and the HCl hydrolysis gives maltose. Invertase and maltase are also being tried for the estn of sucrose and maltose. The saps can also be absorbed on fat free filter paper, extd by 95% alc contg a little NH_4 and the alc ext. treated as above. This modification gives excellent results.

E. J. C.

Effects of hydrogen-ion concentrations on rice cultures. S. K. MITRA AND LOKANATH PHUKAN *Proc 15th Indian Sci Cong* 1928, 166–7 —1 fleets of H ion concns. on rice cultures were studied with Knop's soln. as the standard nutrient soln. and with HCl (0.001–0.01%) and NaOH (0.001–0.05%) as the adjusting media. The pH values of the solns. ranged from 3.0 to about 8.4. The expt was continued with one-month old seedlings for a fortnight and the total length of roots in mm. was taken as the criterion of growth. Results —(1) The seedlings showed strong acid toxicity at a concn of 0.002% HCl, whereas higher concns. were distinctly inhibitory to root growth. With the addn of alkali, the roots developed far better and the highest root development was shown at a concn of 0.3% NaOH. (2) A definite relationship was found between the H ion concns. of the culture solns. and their corresponding root lengths, which showed a steady increase with the higher pH values. pH 3.9 was distinctly toxic whereas pH 3.3 was extremely toxic to root growth. At acidities less than pH 6.0, the development of roots was below normal but beyond this it was quite satisfactory. The highest root growth was attained at pH 7.0, but further on a drop was shown at pH 8.4, which was, however, not quite confirmatory and requires further experimentation.

E. J. C.

Effects of Shive's three salt nutrients on rice seedlings. S. K. MITRA AND LOKANATH PHUKAN *Proc 15th Indian Sci Cong* 1928, 167 —Following Shive's procedure an expt on the effects of the 3 salt nutrients of KH_2PO_4 , $Ca(NO_3)_2$ and $MgSO_4$ was tried on the development of roots in rice seedlings. The nutrient solns. were arranged to give all possible combinations in variations of 0.1–0.8 of the total osmotic concn of 1.75 atm., for each of the 3 salts used. The results of the expt. were as follows: (1) The culture showing the highest root development had 0.5 of its total osmotic concn. due to KH_2PO_4 , 0.2 to $MgSO_4$, and 0.3 to $Ca(NO_3)_2$. Similarly, other high root developments were characterized by high concns. of KH_2PO_4 and low concns. of the other 2 salts. (2) The culture showing the lowest root development had 0.3 of its total concn. due to KH_2PO_4 , 0.5 to $Ca(NO_3)_2$, and 0.2 to $MgSO_4$. The region of low root development was characterized by low concns. of KH_2PO_4 and high concns. of $Ca(NO_3)_2$ or $MgSO_4$, or both. (3) The best and the worst cultures contained the metallic ions of Mg, Ca and K in the proportions of 0.56 : 0.43 : 1.00 and 0.93 : 1.20 : 1.00, resp.

E. J. C.

Injection experiments with special reference to the production of alkaloids, etc., and general metabolism in plants. S. KRISHNA AND H. CHAUDHURI *Proc 15th Indian Sci Cong* 1928, 224 —For exptl purposes 2 or more species (of the same genus) which differed in the production of alkaloids, etc., were taken at a time. The plants were all grown from seeds under observation. Successful and interesting results were

obtained when even min quantities of chemicals in colloidal solns. were injected. Thus when colloidal Fe was injected in opium poppy, which normally contains berberine, it produced less of that alkaloid, whereas when colloidal Fe was injected in red poppy, which contains practically no morphine or berberine, berberine was actually found. Injection of colloidal S caused vigorous growth and produced healthier flowers and seeds, though the flower production was a bit delayed. Injection of colloidal Fe in a no of plants inhibited the development of flower buds. E J C

Importance of iron for the growth of fungi (*Fusarium vasinfectum* Atkinson). B B MUNDKUR *Proc 15th Indian Sci Cong* 1928, 230—*Fusarium vasinfectum* Atkinson was grown in media with and without Fe in order to det. whether Fe exercised any influence on the growth of the fungus. It was observed that in solns. with Fe there was more growth than in solns. without Fe. It is concluded that for some fungi, at any rate, Fe acts as a stimulant, even though they do not have chlorophyll and therefore do not need it. E J C

Studies in photosynthesis in tropical sunlight. I. G GOPALA RAO AND N R DHAR *J Phys Chem* 35, 1418-23(1931)—HCHO has been obtained from CO₂ and H₂O in the presence of the following photosensitizers and sunlight: chlorophyll, methyl ene blue, malachite green, methyl orange, ferric hydroxide sol, uranyl nitrate, Cr sulfate and CuSO₄. Small quantities of CO were obtained from the photochemical reduction of CO₂ by chlorophyll. II. Photosynthesis of formaldehyde and carbohydrates from alkali bicarbonates in the presence of colored insoluble substances. *Ibid* 1424-32.—Exposure of NaHCO₃ solns. contg. morg. catalysts such as Ni carbonate, Co carbonate or colloidal Fe(OH)₃ to sunlight for a period of 45 to 60 hrs. effected the synthesis of definite quantities of HCHO and carbohydrates. HCHO solns. contg. ZnO, methyl orange, or FeCl₃ when exposed to sunlight formed small quantities of reducing sugar. The soln. alone, or contg. Fe(OH)₃ sol. or MgO produced no sugar. A suggested mechanism for the photosynthesis of HCHO consists of 2 steps, (1) the combination of chlorophyll and H₂CO₃ in the presence of light energy to form chlorophyll peroxide and CO and (2) the instantaneous interaction of this nascent CO with H₂O to form HCHO and O₂. The proposed mechanism is shown to account for most of the facts of photosynthesis. P 11 EMMETT

Sorbitol content of Norwegian mountain ash berries. AXEL JERNSTAD *Arch Pharm* 269, 68-9(1931)—Both the acid (malic) and sorbitol contents of Norwegian mountain ash berries increase during the ripening period. In French berries the sorbitol increases while the acid content decreases. The Norwegian berries are lower in sorbitol than the French. W O E

Chemical examination of the red pigments of some autumn fruits. LÁSZLÓ ZECHMEISTER AND LÁSZLÓ CHOLNOKY *Monatsh f. Tiermisch. Chem* 47, 209-17 (German abstract 218)(1930), cf *C A* 24, 3020, 4303—Sixty kg. of *Lycium kalimifolium* gave 17 g. of a cryst. pigment, the compn. of which was found to be C₂₇H₃₄O₄. It is identical with the physalene of Kuhn and Wiegand. No secondary pigments were found. Fruits of *Taraxacum officinale* contained lycopene, also fruits of *Silene acaulis*. Arillus of *Econium europaeum* contained a xanthophyll like pigment of the compn. C₄₀H₅₆O₂. Examin. of other fruits, e. g., *Arum maculatum* and *Sorbus aucuparia* is in progress. S. S. DE FINALLY

Investigations on the action of different ions on the saccharase of *Penicillium glaucum*. ELIA FERÉZ *Metabolismus Kalks* 3, 292-315(1930)—The quantity of saccharase in well nourished *P. glaucum* is almost proportional to the quantity of added sugar. Scarcity of Ca has no influence, but that of Mg or PO₄ ions decreases the saccharase quantity. The optimum μ is 4.5. The no. obtained by multiplying dry-matter content and saccharase quantity of a mold culture is a const. value which is characteristic of the resp. cultures. Dialysis at first increases enzyme action. The action is strengthened by the presence of Cl, SO₄ and NO₃ ions. The enzyme action is, however, diminished by long dialysis and cannot be reactivated with salts. The μ optimum is 4.3 after dialysis. The presence of salts has no influence on this value. The action of saccharase cannot be increased by autolysis. S. S. DE FINALLY

Cyanogenesis in plants. M. E. ROBINSON *Biol. Rev.* 5, 126(1930), *Physiol. Abstracts* 15, 556—A review. P. G.

Vegetable proteins. I. The proteins of *Dolichos lab. lab.* DURAISWAMI NARAYANAMURTI AND CHIMBATORE VENKATARAMANA RAMASWAMI AILAR *Biochem. J.* 24, 1650-24(1930)—The chief protein of *Dolichos L5 L5* constitutes about 80% of the N in the meal and is obtained in the aq. fraction. BENJAMIN HARROW

Irradiation of *Dolichos tyrosinase*. DURAISWAMI NARAYANAMURTI AND CHIMBATORE

RATORY VENKATARAMANA RAMASWAMI AITIAI. *Biochem. J.* 24, 1035-8 (1930).—The enzyme is most active when it is least charged. BENJAMIN HARRISON

Solubilization of leaves. RICHARD HOEMAN. *Univ. Calif. Pub. in Botany* 15, 139-141 (1921).—In *Passiflora mollissima*, solubilization was delayed rather than hastened by an increased CO_2 supply, and occurred more slowly, if at all, when the leaves were washed. JOSEPH S. HARRISON

Fermentation phenomena of the *Xylophaga* cocci and their relation to yeast. J. GRÖN. *Holm. Fysik.* 47, 475-2 (1901).—See C. A. 24, 5588. S. JONES

The enzyme content of dormant and germinating seeds. YUTAKA JOYO. *Ann. Nip. Med. Univ. Tokyo* 13, 211-5 (1931).—The carbohydrate- and protein-splitting enzymes, catalase and diastase, are present in the dormant seeds and become augmented during germination. The lipases are present in wheat seeds, beans and peas and likewise increase during germination. In the hemp seed, watermelon seed and pumpkin seed the lipase is largely demonstrable in either the resting or germinating stages. Barley seeds are free from urase. Peas and hemp seeds manifest only a trace of urase which shows no increase during germination. In the soy bean, watermelon seeds and pumpkin seeds urase is found in considerable quantity which increases during germination. B. S. LAYTON

The role of silica in living organisms. The biology of the diatoms. ERICOUR BACROCH AND M. LEFÈVRE. *J. phys. coll. chim.* 27, 341-9 (1931).—In the presence of ill-defined factors certain marine and fresh water diatoms, grown in sterile culture but in the presence of silica, become incapable of forming a siliceous carapace. No loss of vitality results and the diatoms are able to develop and multiply through many generations. The loss of carapace is more or less rapid according to the species. When a diatom divides in spite of the absence of rigid valves, the resulting forms resemble the original but the death of the species is impossible. The presence of valves is not essential for locomotion. L. A. MARSHALL

Studies on leguminous bacteria and plants. II. Utilization of different nitrogen compounds as well as of the nitrogen accumulated in the root nodules by leguminous plants. ARTHUR I. VERTAN and SYNDYNE VAN HARTEN. *Biochem. Z.* 232, 1-14 (1931), cf. C. A. 24, 3232.—Red clover grows best in media containing amino acids as the N source, and the best results were obtained with hydrolyzed casein. Growth is less with $(\text{NH}_4)_2\text{SO}_4$ and still poorer with KNO_3 . The behavior of white clover toward N compounds is just the reverse, since it grows best with $(\text{NH}_4)_2\text{SO}_4$, and very poorly with amino acids. The amino acids are utilized directly and not as a source of NH_4 , since the pH of the medium remains unchanged and NH_4 cannot be detected. S. MOKETIAN

Biochemical studies in connection with the problem of early budding. ANTONIUS NIETHAMMER. *Biochem. Z.* 232, 1, 1-25 (1931).—Either frost or a dry spell, especially the latter, promotes budding. The effect of frost is to increase the amt. of reducing sugar. The sugar content of a twig in the course of a year is subject to great variations and its max. does not coincide with the max. budding. The lability of the parenchyma cells to plasmolysis seems to correspond to the tendency to budding. S. MOKETIAN

Choline as the injurious substance in diseased barley. G. SCHROETER and L. STRASSBERGER. *Biochem. Z.* 232, 1-2-6 (1931).—It is shown that barley extract with dil. HCl yields lecithin, which on boiling and evapn. of the acid ext. yields choline, so that from every barley and not only from the diseased barley choline can be obtained. Later studies are further encouraged because by his method of pptg. choline from the acid ext. with an alc. soln. of HgCl_2 ketamines and other substances besides choline are separated as complex Hg salts, which were thus erroneously called, as choline. S. MOKETIAN

The death wave in *Nitella*. III. Transmission. W. J. V. OSTERHOUT and S. E. HILL. *J. Gen. Physiol.* 14, 365-92 (1931), cf. C. A. 23, 3004.—When a cell of *N. flexilis* is cut a series of rapid elec. responses are set up, which are transmitted very rapidly. They are followed by slower responses, the speed of which falls off as the distance from the cut increases. The mechanism appears like a mech. disturbance, the intensity of which falls as it travels. The faster responses are probably due to the motion of sap past surfaces which have not been altered. They are possibly similar to elec. changes following a blow on the end of a soft rubber tube containing Ag-AgCl electrodes. The slower responses seem to be due to alterations in the protoplasm and are usually irreversible. C. H. R.

The sensitivity of *Laminaria* to external actions and the volatilization of iodine. PIERRE DANGEARD. *Compt. rend.* 192, 500-1 (1931), cf. C. A. 24, 2160.—The liberation of free I_2 from *Laminaria flexilis* is provoked by a variety of external actions. T. H. RIDER

The constancy of vegetable composition according to Liebig and the sugar beet

bred by selections. *L. MILE SAILLARD* *Compt rend* 192, 504-7 (1931) —Liebig was led to believe that plants had fixed compns. S found that the quantities of basic minerals relative to a given amt of sugar in the sugar beet decreased as the sweetness increased. The H_2PO_4 contained in the entire plant did not bear a const. ratio to the sugar. The relative quantity of base combined with org. acids decreased as the sugar increased.

T. H. RINER

The physiology of latex. MARIA ROBIN *Jahrb. wiss. Botan.* 69, 587-635, *Chem. Zentr.* 1930, 1, 984-90. The methods employed permitted working with very small quantities of latex, thus the latex of the same plant could be exam'd under various conditions. Expts. with dye solns and Li solns showed that the latex tubes are not particularly adapted for the transport of fluids. The solid matter of the latexes varied between 15 and 30% in the euphorbia exam'd. When plants cultivated in moist soil were kept in a dry atm. for 24 hrs. the solid matter of the latexes was reduced to one half, this was probably due to the inhibition of photosynthesis and to the transfer of osmotically active substances to adjacent tissues. Degradation of starch was observed in the latex of plants kept in dry air. Latex of plants treated with $CuCl_2$ soln showed degradation of starch and mercuric in reducing substances. Inhibition of assimilation by darkening also caused a diminution in solid matter. Treatment of cut off shoots with salt solns produced a distinct decrease in starch, a pronounced increase in sugar and an increase of solid matter. Glucose soln (0.4%) caused a decrease in solid matter. The electrolyte content increases after administration of sugar, which may be due to the formation of acid. The electrolyte content is of the same order of magnitude in the representatives of the different genera and species of the Euphorbiaceae, Moraceae and Compositae, in a diln. 1:1000 the sp. cond. ranged from 10×10^{-4} to 30×10^{-4} . Cultivating a plant in a dry atm. causes a reduction in cond., which is compensated by transferring the plant to a moist atm. Treatment of cut off shoots with $CuCl_2$ soln resulted in a strong increase in cond. The total press. juice had a considerably smaller cond. than the latex of the same plant, with respect to the order of magnitude it is equiv. to the press. juice of the latex free plants from ordinary habitats. The latex has a storing function, this fact does not imply that the substances stored are permanently eliminated from the metabolism. The latex tube system has no conducting function, but it seems to be of importance for the regulation of osmosis.

G. SCHWEN

The carbon dioxide assimilation of arctic plants and the dependence of the assimilation on the temperature. D. MÖLLER *Planta* (Aht. F., *Z. wiss. Biol.*) 6, 22-39, *Chem. Zentr.* 1929, II, 2570. —In Dadravn, West Greenland, the assimilation of 2 arctic plants, *Salix glauca* L. and *Chamaenerium latifolium* (L.) Spach, was exam'd in the month of July. From the curves on assimilation and illumination intensity it is found that at midnight in the long day period *Chamaenerium* has a weakly pos. apparent assimilation at temps. of below 10° . The assimilation curves of arctic plants show the same course as those of the heliophilous leaves of Danish plants. With completely open stomata and normal CO_2 content of the air the assimilation was exam'd. at different illumination intensities at 10° and at 20° . The compensation point and the respiratory intensity increase with the temp. With low illumination intensities the apparent assimilation is greater at 10° than at 20° , with illumination intensities larger than 4000 or 6000 lux, resp., it is greater at 20° than at 10° . The course of the curves makes Lundegårdh's temp. assimilation curves appear improbable. The stomata of arctic plants are rather wide open during the whole night in some cases up to 100%. The quantity of the lignified arctic plants per area is lower than that of the subarctic plants, the same is true for the leaf area per surface unit.

G. SCHWEN

Differential growth of phytophthoras under the action of malachite green. LEON H. LEONIAN *Am. J. Botany* 17, 671-7 (1930) —Malachite green, at the rate of 1 part dye to 1, 2, 4, 8 and 16 million parts of nutrient soln, was used to induce differential growth in *Phytophthora*. Three organisms failed to grow when malachite green was present at the rate of 1 part dye to 16 million parts of nutrient soln, and 8 others made only a sporadic growth. Twenty-one organisms failed to grow when the amt. of dye was doubled, and 29 others made only a sporadic growth. Only 1 organism was able to make a growth in nutrient soln with 1 p. p. m. of malachite green. The critical concn. of the dye for the *P. cactorum* group is in the neighborhood of 1 part malachite green in 4 million parts nutrient soln, while that for *P. omanora* is decidedly higher.

J. J. SKINNER

Effect of mineral salts upon the transpiration and water requirement of the cotton plant. BERNARD S. MEYER *Am. J. Botany* 18, 79-83 (1931) —The salts employed for the transpiration expts. with cotton in soil were $NaCl$, $NaNO_3$, KCl , KNO_3 , $CaCl_2$ and $Ca(NO_3)_2$. The first 4 were applied to the soil in concns. of 0.025, 0.05, 0.1, 0.2 and

0.4% of the soil, with the Ca salts an addnl concn of 0.8% was used. For 4-day periods addition of any of these salts in all the concns employed, with the exception of the 2 lowest concns of KNO_3 , resulted in a decrease in transpiration as compared with plants growing in the soil to which no salt was added. Transpiration progressively decreased in amt. with progressive increases in the concn of any of these salts in the soil. The results were essentially the same whether transpiration was calcd. on the basis of leaf area, fresh wt. of the top or dry wt. of the top. It is clear that the concns of salts employed with reference to the soil conditions used lie in the range where osmotic effects predominate over specific ionic effects. NaCl in concns of 0.0125, 0.025 and 0.05% of the dry wt. of the soil, and CaCl_2 in concns of 0.0125, 0.025, 0.05, 0.1 and 0.2% of the dry wt. of the soil, were used in the water requirement study. Cotton plants were allowed to grow under these treatments for 50 days, the usual technique of water requirement studies being followed. The water requirement of the plants treated with NaCl decreased progressively with increased concn of this salt in the soil. The 2 lowest concns of CaCl_2 employed resulted in an increased water requirement; beyond that concn the water requirement decreased progressively with increase in the concn of CaCl_2 in the soil. A progressive increase in the concn of NaCl or CaCl_2 in the soil resulted in a progressive increase in the final water content of the tops of the cotton plants.

J. J. SKIVER

Chemistry of the halophytes. II. J. ZELLNER AND E. ZIRMUNDA. *Monatsh* 56, 197-9 (1930), cf. *C. A.* 21, 2718.—The ash content of *Lepidium crassifolium* L. was 10.35%. It had the following compn: K, 10.2%, Na, 21.4%, and Mg, 3.46%. The following substances were isolated from *Salicornia herbacea* L. by the use of selective solvents: an unidentified wax alc., m. 64°, a sterol of higher m. p.; resin and fatty acids which were not identified, choline, phlobaphenes, betaine, Au salt, m. 197-8°, a 1 rotatory invert sugar, phenyllosazone, m. 263°, oxalic acid, and a considerable amt. of inorg. salts.

C. W. SONDERY

Chemistry of the higher fungi. XXI. *Polyporus sulfureus* L. and *Lentinus squamosus* Schroet. JULIUS ZELLNER AND EUGEN ZIRMUNDA. *Monatsh* 56, 200-3 (1930), cf. *C. A.* 23, 3005-6.—From the first named fungus were isolated: an unidentified crystalline substance m. 25°, a cerebrin like substance, m. 112-4°, ergosterol, m. 163-5°, fungi-sterol m. 146°, stearic and palmitic acids, choline, fumaric acid, m. 284°, mannitol, m. 165°, and K II oxalate. The last named fungus yielded mannitol; choline, a phlobaphene like substance, a sugar, phenylglucosazone, m. 206°, and a polysaccharide which when hydrolyzed yielded only glucose. No pentoses were obtained.

C. W. SONDERY

Comparative plant chemistry. XXII. Chemistry of barks. 7. NORBERT FROMMELT, J. ZELLNER AND F. ZIRMUNDA. *Monatsh* 56, 204-11 (1930), cf. *C. A.* 23, 3005.—From the bark of the branches of *Morus nigra* L. there was isolated a phytosterol m. 122°, ceryl alc., stearic, myristic, and linolenic acids, H_3PO_4 , glycerol, tannins, phlobaphenes and a sugar, phenyllosazone, m. 204°. The following substances were isolated from *Alnus incana* L.: protanulin, m. 276°, alnulin, m. 270°, alnuresinol, m. 192-3°, an unidentified substance m. 194-5°, Ac deriv., m. 253°, resin acids and phlobaphenes and a sterol like substance, m. 270°, Ac deriv. m. 231°.

C. W. SONDERY

The absorption of iodine by the roots of plants. J. STOKLASA. *Protoplasma* 8, 193-214 (1929).—Larger amts of I both org. and inorg. accumulated in the organs of plants when the soil contained 1000% of I per kg. than when the soil contained only 20-400% per kg. When the I content of the soil was below 400% per kg. the absorbed I₂ was converted into an org. form. Plants with a low pH were more sensitive to the toxic effect of I.

M. H. SOTLE

Hydrogen-ion phenomena in plants. I. Hydron concentrations and buffers in the fungi. J. I. ARMSTRONG. *Protoplasma* 8, 222-69 (1929).—The tissue and expressed sap reactions of 10 species of the larger Basidiomycetes were investigated. For *Hypholoma fasciculare* the inorg. phosphates of the sap accounted for 50% of its total buffering at pH 6.0-9.0. With *Coprinus miscaneus* the inorg. phosphate in the sap accounted for all the buffering at pH 6.0-7.0, the oxalate at pH 4.0-5.0. With *Collybia velutipes*, phosphate, citrates and malates accounted for the buffer curve. II. An investigation of the buffer complex of sap from stems of *Pelargonium* sp. *Ibid.* 313-43.—The aerial stems of *Pelargonium* from young and mature plants were cut into short lengths and the sap was expressed by pressure. The juices were passed through Gooch crucibles and the filtrates were used at once for titration or analysis. Phosphate and the characteristic odoriferous acid of the tissues, pelargonic acid, were shown to be unimportant constituents of the buffer complex. Org. acids constituted the major part of the buffer complex at the initial pH of the sap. The acids found and estd. included oxalic, malic,

citric and tartaric, and the contributions made by all these buffer systems to the buffering exhibited by the sap were considered in detail. III. The acidity of certain cell walls considered in relation to the higher fatty acids. *Ibid* 508-21(1930)—Higher fatty acids belonging to the acetic and oleic series together with certain derivs of oleic acid, notably ricinoleic or suberic acid, were tested with a series of 9 indicators that had been used in this series of studies. The virage obtained from any given acid with any given indicator was found to depend upon the avidity of the acid concerned for the indicator base. The base avidities of the acids varied in a peculiar manner with the no of C atoms in the acid mol. M H SOULE

The accumulation of ions in living cells—a non-equilibrium condition. S C BROOKS *Protoplasma* 8, 389-412(1929)—Living cells of *Valonia macrophyssa* Kütz were collected from their natural habitat and exposed for different lengths of time to sea water in which the proportion of K had been increased or decreased by the addn of isotonic KCl or NaCl solns. The concn of K in the sap ordinarily increased more than that of Na or Cl regardless of whether it was present in more or less than normal concn in the surrounding soln. The Cl concn in the sap ordinarily changed in much the same way as the K concn but to a significantly smaller extent. The Na concn appeared to have increased first and then decreased, depending upon the length of exposure and the extent to which the sea water was altered. This was explained by supposing (a) that the cell is normally in the non-equal condition with respect to the surrounding soln., (b) that it is surrounded by a plasma membrane consisting of a mosaic of anion permeable and cation permeable areas which are of the nature of charged porous films, and as such exaggerate differences between the diffusion velocities of the ions to which they are permeable, (c) that the diams of the pores are variable according to exptl conditions, (d) that the penetrabilities of different ions are characteristic functions of their own effective diams and of the diams of the porous membranes. M H SOULE

The influence of frequency of cutting on the productivity, botanical and chemical composition and the nutritive value of "natural" pastures in Southern Australia. J GRIFFITHS DAVIES AND A H SIM *Council Sci Ind Research (Australia) Pamphlet* No 18, 5-28(1931), cf Shutt, Hamilton and Selwyn, *C A* 24, 5792—Five series of pastures, approx 40 years old, were compared, the 1st 4 being cut every 2, 4, 7 and 10 weeks, resp., and the last was cut only at the end of the season, i e, November. The yield of dry matter and of crude fiber from the whole pasture was greatest in No 5, and least in No 1, the yield of total N and of crude protein varying inversely. However, though this held for the erect species which comprises most of the plants, *Erodium botrys* gave a max yield of dry matter in No 1 and a min in No 5. The yields of CaO and of P_2O_5 from the whole pasture were not greatly affected by the defoliation, but P_2O_5 was distinctly least in No 5. The CaO content of all the species was high, about 2.2% of the dry matter. The yield in series 1 was greatly affected by rainfall, particularly at the end of the season, the yield of any cut being influenced principally after a lag of 2 weeks. The effect of temp was less marked, but may be a limiting factor near the beginning and end of the season. No significant changes in chem compn of the pasture of No 1 were observed throughout the season. Analyses of the soil and botanical compn. of the various cuts over 2 seasons are given. K V THIMANN

Movement of organic materials in plants. ALDEN S. CRAFTS *Plant Physiology* 6, 1-14(1931)—C suggests that the movement of sugars, amino acids, etc., in plants takes place in the phloem as the channel for conduction with the majority of transport taking place in the cell walls. The functioning of the mechanism proposed is discussed in detail and expts are described in support of C's views. WALTER THOMAS

Food reserves in relation to other factors limiting the growth of grasses. L. F. GRABER *Plant Physiology* 6, 43-71(1931), cf *C A* 22, 1177—Certain cultural practices with grasses (blue grass, red top, fescue and timothy) are correlated with the utilization and accumulation of labile org substances—the so-called org food reserves of plants. The productivity, as measured by both root and top growth, varied inversely with the frequency of defoliation. Frequent and close cutting of the succulent top growth of grasses having abundant reserves resulted in heavy depletion of the available N of the soil. The agronomical aspects of the morphological and ecological relationships of the org food reserves of the flora of grass lands are also discussed. WALTER THOMAS

Relation of hydrogen-ion concentration of tissue fluids to the distribution of iron in plants. R. A INGALLS AND J W SHIVE *Plant Physiology* 6, 103-25(1931), cf *C A* 23, 865—The pH of the tissue fluids of buckwheat, clover, sedum and bryophyllum varied directly with the change of light intensity from day to night (*C A* 19, 3293). The range of variation was proportional to the degree of succulence of the plants.

The p_n of the tissue fluids of the stems of buckwheat and clover was always higher than that of the leaves, but in sedum and bryophyllum the differences were small. In the above plants and also in tobacco, tomato, asparagus and soy bean the sol (filterable) Fe varied inversely and total Fe directly with the changes in p_n produced by variation in the light intensity. The detn of p_n has no significance unless made on material collected at the same time and grown under approx identical conditions. The paper is illustrated with 10 figures.

WALTER THOMAS

The Corail pea or *Adenanthera pavonina*. J. PIERAERTS, E. CASTAGNE and L. ADRIAENS. *Mat grasses* 22, 875-876, 12, 883-884, 886-88, 889-90, 892-4, 897-8, 900-3-4, 903-4 (1930).—The seed of this pea contains neither alkaloids nor tannin. Most of the K is found in the spermoderm and in the albumen while most of the P is in the embryo as lecithin. The albumen is composed mostly of galacto mannans. The sugars, mostly found in the embryo, are a mixt of saccharose and stachyose. The fats of the seeds are the glycerides of oleic, linolic and lignoceric acids. The fat accounts for 75% of the total solid acids. Other acids are found in very small quantities. From its characteristics the oil of the Corail pea must be classed as a vegetable grease. However, only the butter of *Parkia africana* possesses similar const. The const. of *Parkia* butter, Corail pea butter from India and Corail pea butter from Belgian Congo are, resp. acidity index 2.5, 0.56, 1.25, sapon index 154.5, 151.4, 176.6, I index 91.1, 87.9, 94. Reichert Meissl index 0.6, 1.22, 0.77, Hehner index 95.5%, 95.5%, 95.3%. n_D^{20} 1.470, n_D^{25} 1.4570, n_D^{25} 1.4710.

P. THOMASSET

Coloring matter of awobana (Kuroda) 10. Thevetin, a crystalline glucoside from the seeds of *Thevetia nerifolia* (Ait.) 10.

E—NUTRITION

PHILIP B. HAWK

Sun rays and vitamin D. A. VAN WYK, F. H. REERINK and W. MÖRIKOFFER. *Strahlentherapie* 39, No. 1, 80-92 (1930).—The authors compared the action of rays from the sun and those from a Hg quartz lamp upon ergosterol soln. in hexane and found 2 essential differences, the radiation of the ergosterol soln. in the sun rays caused a quicker destruction of the formed vitamin D and also the formation of another decompn. product of unknown compn. The different distributions of the intensities of the radiation of the 2 sources of light are believed to be the explanation for the differences. They found that the relation for vitamin destruction and vitamin formation is 3 times as large with the sun rays as with the Hg quartz lamp. Therefore, with sun radiation much quicker decompn. of vitamin D will take place so that only a small portion of ergosterol is transformed into vitamin D, while a considerable portion of the latter is decompd.

F. R. GREENBAUM

Blood studies in hemorrhagic anemia. H. S. MATHERSON and HENRY LAUREN. *J. Nutrition* 3, 451-63 (1931).—Severe secondary anemia was produced and maintained in 27 dogs by the method of Whipple and Rabschert Robbins (*C. A.* 19, 2692). In dogs fed the standard bread S diet (W and R R., loc. cit.), contg. salmon, the av. weekly production of hemoglobin was 4 g. above the maintenance level. With dried apricots the value was 15, dried peaches 7 and lettuce 10. After several weeks of bread S feeding there is a gradual rise in the no. of red blood cells accompanied by a decrease in corpuscular vol. and corpuscular hemoglobin. The percentage of hemoglobin in these smaller cells remains relatively unchanged. If lettuce, apricots or peaches are added to the diet, these changes do not occur and they are interpreted as being compensatory reactions to the O_2 deficiency occasioned by the low hemoglobin when the severe anemic level is maintained.

C. R. FELLERS

The incorporation of vitamins in bread. F. F. TISDALL, T. G. H. DRAKE and ALAN BROWN. *Can. Med. Assoc. J.* 24, 210-3 (1931); *C. A.* 24, 2161.—Bread is a particularly suitable carrier for vitamins on account of its universal distribution and large general consumption. The refined grain products are deficient. Incorporation of wheat germ will supply B₁ and F, and of irradiated ergosterol, D. Such procedures can be utilized at no addnl. cost to the consumer.

A. T. CAMERON

Microbiological studies as a basis for vitamin investigation. WERNER KOLLATH. *Wiener klin. Wochschr.* 44, 277-80, 320-3 (1931).—A review.

D. B. DILL

Influence of E-vitamin on hypertrophy of female genitalia. F. VERZAR. *Debreceni Tiszt. Tud. Tars. II. Orv. Munk. 3*, 352 (1930).—Lack of E vitamin caused sterility in rats. Hypertrophy of the uterus was caused in young animals of 50 g. wt. by m-

jecting at intervals of 1-2 days 0.5-1.0 cc. concd ext. of E vitamin without sitosterol

S. S. DE FIVELY

Practical significance of alkaline earths in the ash of forages. L. URBANEK *Mezőgazdasági Kutatások* 2, 411-51 (1929), cf. C. A. 24, 4325.—Too much CaCO_3 has a toxic action on sheep and pigs. Salt metabolism is disturbed if forage ash contains more than 30 mg equiv. of alk. earths.

S. S. DE FIVELY

Role of photoactivity in the detection of the antirachitic effect of substances. L. URBANEK *Mezőgazdasági Kutatások* 3, 491-6 (1930).—No connection was found between the antirachitic effect and photochemical activity. Many org. solvents showed photoactivity after a previous treatment with ultra violet light, e. g., benzene, toluene, xylene, nitrobenzene, nitrophenol, aniline, ethyl, propyl and amyl alcs., CHCl_3 , CCl_4 , oleic acid, turpentine oil, paraffin oil and cod liver oil.

S. S. DE FIVELY

The action of white phosphorus and of vitasterol D (Vigantol) on the respiratory metabolism of rachitic young rats. H. SLERL *Arch. exp. Path. Pharmacol.* 140, 194-204 (1929). *Physiol. Abstracts* 15, 373.—A rachitogenic diet depressed the metabolism of rats. P injected subcutaneously in oily soln. in doses of 10^{-2} to 10^{-4} mg. daily caused the metabolism to rise to normal, but did not greatly increase the wt. or cure the rickets. Irradiated ergosterol caused a rise of metabolism and cured the rickets.

G. G.

Protein diet and the number of leucocytes in the intestinal mucosa of the mouse. H. DE WINTWARTER *Compt. rend. soc. biol.* 102, 121-3 (1930). *Physiol. Abstracts* 15, 368.—An exclusively albuminous diet causes a characteristic invasion of macrophages in the mucosa, which modifies the villous formations. This is accompanied by a considerable relative increase of eosinophiles and intense destruction of leucocytes, the debris of which is found in various coats of the intestine. No similar results are seen on a purely fatty or carbohydrate diet.

G. G.

Determination of vitamin A. KATHARINE H. COWARD, KATHLEEN M. KEY, FREDERICK J. DAVER AND BARBARA G. E. MORGAN *Biochem. J.* 24, 1932-60 (1930).—The vitamin A contents of 2 substances are compared, 2 groups of rats (5-10 in each group) being used, each group is fed on A free diet until growth ceases, and each rat within the group is then given the same daily dose of the substance to be tested. The mean increases in wt. of the 2 groups are noted. These mean increases are then referred to a curve which relates mean increase to dose of a particular sample of cod-liver oil.

BENJAMIN HARROW

Salicylic acid fruit in the prevention and treatment of rheumatism in children. J. FRSTEN *Arch. Pediatrics* 48, 73-81 (1931).—Salicylic, benzoic, citric and tartaric acids are antirheumatic. For this reason, certain fruits (strawberries, huckleberries, raspberries, plums, cherries, lemons, grape fruit and melons) should be included in the diet beginning with the second year of life, both as preventives and as a portion of the therapy of rheumatic fever.

JOSEPH S. HERRMAN

Low-fat, high-starch evaporated-milk feeding for the marasmic baby. JOHN HOWLAND WEST *Arch. Pediatrics* 48, 189-93 (1931).—The marasmic baby usually does not tolerate fats and sugars. A readily digested, high-calorie diet, composed of unsweetened evapd milk, Ca caseinate and cooked starch, usually is satisfactory.

JOSEPH S. HERRMAN

To better human life by milk. I. H. W. A. PRICE. *Creamery and Milk Plant Monthly* 20, No. 3, 26-37, No. 4, 46-51 (1931).—Data obtained indicate that the vitamin content of dairy products is largely dependent upon the utilization by the cow of a rapidly growing fresh grass that has developed on a soil amply provided with such mineral requirements of the plant as to provide the cow not only with these minerals but with activating substances which are not produced in adequate quantity in the absence of these minerals. Dets. of vitamins A and D were made on butters obtained from various parts of the world. Seasonal variations were found in the quantities of these vitamins present in the butter. Thus in North America vitamin D tends to be low from Nov. to April, rising in May and June, falling in midsummer, sometimes rising in the early autumn with a rapid decline in the fall and early winter. It was found that the vitamin content of butter did not decrease during storage. From the standpoint of public health, therefore, it was recommended that butter should be placed in storage at the season of its greatest vitamin content.

A. H. JOHNSON

Observation on the assay of the antineuritic vitamin. Some of the factors involved in the use of the rat method. W. H. SEBRILL AND E. ELYOVE *U. S. Pub. Health Repts.* 46, 917-25 (1931).—Results are reported which are in agreement with the experience of others, to the effect that the symptoms of polyneuritis in rats appear to be associated with shortage rather than complete absence of antineuritic vitamin. The

cureative method for testing antineoplastic concentrates on rats may be applied by injecting a suitable soln. of the concentrate, subcutaneously or intraperitoneally. J. A. K.

Nutritive value of soy-bean cake for hens. II. Kozo SUGIYAMA AND TADASHI HAYASHI. *J. Agr. Chem. Soc. Japan* 6, 931-9 (1931).—Soy-bean oil cake was given as protein source of the feed. Twelve parts of bone powder, 4 parts of CaCO_3 and 4 parts of NaCl for 100 parts of the cake were supplied. The nutritive value was similar to that of fish meal. It gives results on egg production and wt. of egg similar to those of other animal feeds. III. Fed 910-6.—The chicks hatched from the eggs as above mentioned were also fed with soy bean oil cake. No abnormal signs were noted. Y. KIHARA

Vitamin D. III. ETSUO TAKAMURA. *J. Agr. Chem. Soc. Japan* 6, 977-74 (1931), cf. C. A. 25, 537.—Ergosterol was dissolved in olive oil and treated with ozone. The ergosterol soln. so treated was antirachitic for mice. Excess of ozonated diminished the antirachitic activity. If ozone acted on ergosterol for 3 min. rickets were cured by a dose of 0.03-0.1 mg. ergosterol per day. Vitamin D itself may be Δ^7 -cholecalciferol. Y. KIHARA

Soy-bean cake for the fattening of swine. Kozo SUGIYAMA. *J. Agr. Chem. Soc. Japan* 6, 975-8 (1931).—Soy bean oil cake as 20 and 30% was added to a feed consisting of korean 47, maize 31, bone powder 7 and NaCl 1%. The nutritive value of the feed with soy bean oil cake was as good as that with fish meal. The results in fattening were rather superior. Vitamin A, Cl, Na and Ca should be supplied. Y. KIHARA

Muramatsu's beriberi bacteria and antismore. TOSIYU ARITAMA. *J. Agr. Chem. Soc. Japan* 6, 1123-43 (1931).—Mice fed with polished rice excrete Muramatsu's beriberi bacteria. If the parents are fed with a diet deficient in vitamin B, the bacteria are difficult to find in the excreta. Mice fed with polished rice show the symptom of the sickness but do not excrete the bacteria. The bacteria could not be found in the excreta of mice fed with polished rice. There is no relation between beriberi and Muramatsu's bacteria. Y. KIHARA

Effects of a large quantity of lactose and in the diet. RYUZO SASAKI. *J. Agr. Chem. Soc. Japan* 6, 1144-52 (1931).—Mice were forced to eat lactose and as 5.5% of the diet. Ca and especially P of the bone were decreased. Decrease of P in the muscle was also found but it was less than that of the Ca decrease. Y. KIHARA

Vitamin C. IV. Vitamin C in seeds germinated under a Mazda lamp. TOMIJI MATSUKA. *J. Agr. Chem. Soc. Japan* 7, 14-28 (1931), cf. C. A. 24, 5349.—The seed-rice was germinated under 1000 w. elec. lamp at 28-31°. Vitamin C was produced after 3 days of germination. The detection of vitamin C by Bersonson's reaction does not coincide with the biological test. Y. KIHARA

Effects of alkali digestion upon the vitamin content of cereal straws. RYUZO IWATA. *J. Soc. Agr. Sci. (Japan)* No. 325, 37-45 (1931).—The straws of barley, wheat and rice contain some vitamins A and B. The vitamins in the straws were more stable toward 1% lime, especially in the soaking process, than toward NaOH . Y. K.

The study of calcium metabolism. Citrate and urinary excretion of calcium. R. PROVERMAN AND L. BRILL. *Bull. soc. chim. ind.* 12, 1151-7 (1931).—Subcutaneous injection or feeding of tri-Na citrate caused a marked increase in urinary Ca and a decrease in blood Ca. The presence of citrate in body tissues and fluids is thought to have an important relation to Ca metabolism. C. G. KING

Loss of vitamin A during the baking of thin butter cookies. HELEN T. PARSONS, IVA STEVENSON, IVA MULLEN AND CAROLYN HORN. *J. Home Econ.* 22, 566-72 (1931).—Thin butter cookies, baked at 160° for 10 min., lose 23-25% of their original vitamin A content. AMY LEVINSOHN

Metallized food in the regeneration of hemoglobin in rat and man. J. L. MCGHEE. *Science* 73, 247-8 (1931).—Milk was left in contact with an alloy of Fe, Co, Mn and Cu in an ice box and small quantities of these metals were dissolved in it. This metallized milk and control milk were fed to young white rats whose hemoglobin had been reduced to about 75%. Those to which metallized milk was fed showed rapid regeneration. They also showed some storage of these metals. Cu alone was effective but the mixt. of metals was much better. There was a lower mortality among rats fed on metallized milk than on milk to which salts of the metals were added. Tests extended to man indicated similar results. G. H. W. LUCAS

The comparative value of irradiated ergosterol and cod-liver oil as a prophylactic antirachitic agent when given in equivalent dosage according to rat units of vitamin D. D. J. BARNES, M. J. BRADY AND E. M. JAMES. *Am. J. Diseases Children* 39, 45-58 (1930).—Cod-liver oil, administered to infants in doses of 1400 rat units of vitamin

a seasonal variation in litter size, being higher in summer. The omnivorous mothers were heavier at parturition than the vegetarian, but for mothers of the same body weight the young from the omnivorous animals were heavier. The birth weight of the young tended to increase with the age of the mother and the order of the litter up to the fourth for the vegetarian and to the third for the omnivorous rats. L. A. M.

Growth of vegetarian rats on omnivorous diet. SHING WAN and TUNG-TOU CHEN. *Chinese J. Physiol.* 5, 71-8(1931)—Young born of vegetarian mothers, but nursed by omnivorous mothers and reared upon an omnivorous diet after weaning, averaged 83% normal in wt. at 4 months. Young born of and nursed by vegetarian mothers, and placed on an omnivorous diet at 4 weeks, attained wts. which were 82% normal. Similar young reared upon a vegetarian diet to 4 months and then changed to the omnivorous diet reached wts. which were 77% normal. "These results show that nutrition during lactation is more important than later in life." L. A. MAYNARD

Excessive doses of vitamin D. M. FISCHER and V. NAVA. *Riv. clin. pediatrica* 28, No. 3(1930); *Rev. sud americana endocrinol. inmunol. ginecologica* 14, 27(1931)—High doses of vitamin D produce in rats alterations in the bones, the blood vessels, kidneys and spleen. The distance of the toxic dose from the therapeutic is great enough to avoid danger. A. E. MEYER

The toxic and calcifying action of irradiated ergosterol. L. MICHELLE. *Arch. ital. biol.* 84, 111-7(1931)—Injection of large doses of irradiated ergosterol into rabbits causes pronounced cachexia and deposition of Ca salts in such organs and tissues as normally exhibit no calcification. There is also evidence of a hemolytic effect. S. MORGLIS

The action and the mechanism of the action of excessive doses of irradiated ergosterol in experimental parathyroid insufficiency. M. COMEL. *Arch. ital. biol.* 84, 114-20(1931)—Large doses of irradiated ergosterol administered previous to parathyroidectomy prevent the occurrence of tetany and of the symptoms of parathyroid insufficiency. Irradiated ergosterol in the amts. used causes considerable hypercalcemia and also a slight hyperphosphatemia, but the hypercalcemia does not last long following parathyroidectomy; after a variable interval it gives place to hypocalcemia characteristic of parathyroid insufficiency. However, the blood P does not increase as much as it usually does in the parathyroidectomized animal so that the Ca/P ratio does not fall nearly as low. In animals treated with the ergosterol after the operation the hypercalcemia is transitory, and is followed by hypocalcemia. Also in *Arch. ital. biol.* 29, 123-44(1930). S. MORGLIS

The bread problem. II. The physiological effects of whole-grain bread. I. ABELIN. *Biochem. Z.* 232, 278-91(1931), cf. C. A. 24, 1145—Unlike white bread or bread rich in bran, bread made from 100% whole grain flour (as 75% whole wheat and 25% whole rye) has good nutritive value. The value of this bread is still further enhanced when the excess H_2PO_4 of the flour is largely neutralized by means of Ca salts of org. acids. Young rats nourished exclusively on pure whole grain bread or on one contg. Ca attain their normal body wt. and are in good health even after 6 months on such a regime. Fed on white bread or on bran bread, the animals grow very slowly, show a disposition to skin and mucous-membrane infections and generally die in 3-4 months. Neutralizing the excess H_2PO_4 of white bread with Ca had some favorable effect but not nearly so much as in the case of the whole grain bread. Whole grain flour may, of course, be of good or poor quality, and the method of prepg. the bread is also important in detg. its nutritive value. For instance, toasting made the bread less nourishing and even injurious. Animals fed on toasted whole grain bread manifested bone injuries, became very sick and died. S. MORGLIS

Ammonium bicarbonate together with acid sugar-beet cossettes as a protein substitute. (A feeding experiment on milk goats.) FRANZ-HERMANN ZIEGLER. *Biochem. Z.* 232, 323-42(1931)— NH_4HCO_3 could replace about 50% of the protein in a ration designed to maintain an increased milk production in goats. Sixty five tables of detailed results are given. S. MORGLIS

The nitrogen balance during dietary corrections of obesity. J. M. STRANG, H. B. McCLUGAGE and FRANK A. EVANS. *Am. J. Med. Sci.* 181, 346-49(1931)—The level of N metabolism in obese patients does not differ from that of normals. The rapid reduction of wt. by a "maintenance protein diet" produces only a slight depression in the N level even after 28 weeks. The total N loss in the period of neg. balance is less than 4% of total body N. R. C. WILLSON

Creatinine excretion in abnormal states of nutrition. H. B. McCLUGAGE, GEORGE BOOTH and FRANK A. EVANS. *Am. J. Med. Sci.* 181, 349-55(1931)—The creatinine excretion of obese subjects is close to that of normal ones. The excreted creatinine

does not change appreciably when reduced by dietary measures alone. The excretion in persons of subnormal wt. is markedly reduced as compared to the normal.

R. C. WILLSON

The effect of irradiated ergosterol on the composition of gastric and pancreatic juices. WALTER BAUER, ALEXANDER MARBLE, STEPHEN J. MAGDOCK AND JOSEPHINE C. WOOD. *Am J Med Sci* 181, 399-413(1931).—1 rec HCl was reduced in 3 of 4 patients, a questionable reduction occurring in the fourth case. The pancreatic enzyme activity was depressed in 1 case. The fixed base content of pancreatic juice was not affected. Cl ion was decreased and CO_3 ion increased and thus a more alk. pancreatic juice resulted.

R. C. WILLSON

The influence of vitamin A on the action of viosterol in rickets. P. ROIMER AND R. DUBOIS. *Rec franc pediatrie* 6, 604-8(1930), *J Am Med Assoc* 96, 305.—Six rachitic children were given either the min effective dose or an insufficient dose of viosterol (from 300 to 600 rat units per day). They were also given daily doses of 3000 rat units of vitamin A. The addn of the larger doses of vitamin A did not modify the action of the viosterol in any way. The clinical symptoms, results of x ray examn and blood analysis corresponded with those normally obtained by small doses of viosterol.

R. C. WILLSON

LEICH, J. NEIL. Dietetics in Warm Climates Including Food-stuffs, their Analysis and Role in Disease. London: Harrison and Sons, Ltd. 486 pp. 25s., net. Reviewed in *Indian Med Gaz* 66, 228(1931).

MELLANDY, MAY. Diet and the Teeth. An Experimental Study. Part II. A Diet and Dental Disease. B. Diet and Dental Structure in Mammals other than the Dog. London: H. M. Stationery Office. 2s. 6d. Reviewed in *Pharm. J* 126, 386(1931).

Antirachitic products. SOC. DES USINES CHIM. RHÔNE-POULENC. Fr. 698,040, June 11, 1930.—Products having high antirachitic activity are prepd by irradiation of ergosterol with ultra violet rays. The ergosterol is sepd from the crude product of irradiation by using the low soly of ergosterol in org solvents such as EtOH, acetone or AcOEt. The sepn of the ergosterol is followed up until a dextrorotatory product is obtained, which has preferably a rotatory power in alc soln above 25° for the yellow Hg ray. The irradiation of the ergosterol is stopped before the point of max. activity is passed and preferably before it is reached.

F—PHYSIOLOGY

E. K. MARSHALL, JR.

Excitability of the reflexes as a function of the p_{H} value. M. MIROLS. *Att accad. Lincei* 12, 191-4(1930), cf *C A* 23, 3482, 4253.—The effect of changing p_{H} on the cerebrospinal axis of *Bufo vulgaris* was studied by immersing this for half hr periods in acid solns (HCl, lactic acid), as well as in alk. solns (NaOH, KOH) and noting the reaction. Outside of the limits p_{H} 6.7-7.2, either acids or alkalis promote more and more excitability increasing to clonic convulsions, tetany, and finally total paralysis at values below p_{H} 6.1 or above 7.6.

A. W. CONTIERI

The regulating function of the central nervous system. M. MIROLO. *Att accad. Lincei* 12, 246-51(1930).—The function of the central nervous system of *Bufo vulgaris* in regulating p_{H} value has been studied by adding small amts (in mg.) of the substance to 10 cc of the acid mixt, Sorensen's solns being used (0.1 M HCl + 0.1 M glycocoll, M/15 K_2HPO_4 + M/15 Na_2HPO_4) as well as those of Mellvaine (0.1 M citric acid + 0.2 M Na_2HPO_4). Various parts of the nervous system, as brain and spinal fluid, have the same regulating function. Time of contact or temp has little effect on the regulating function, there is no difference in the action of isotonic and anisotonic solns, however, in isotonic solns of mixts of KCl, LiCl, KIO_3 , slightly greater regulation is obtained than in isotonic NaCl alone. A neutral strychnine sulfate soln. (2.5 10,000, tetany-producing) raises the power slightly, whereas a soln of 9.5 cc physiol soln and 0.5 cc abs EtOH (narcotic) depresses the action. A. W. C.

The hydremic curve (fasting) in various short exercises. III. L. BRACALONI. *Att accad. Lincei* 12, 252-7(1930), cf *C A* 24, 4314.—Previous work suggested that the hydremic curve may vary with the intensity of exercise rather than duration, therefore, 6 subjects were subjected to various forms of exercise: 50 and 100 m race, climbing a 9 m rope, rowing 10, 20, 30 strokes, 10-30 cunnings, 10 leg flexings. In all cases the water content of the blood diminished 0.7 to 1.23%, but was recovered in about 15 min, and differed more with the individual than with the form of exercise. A. W. C.

New viewpoints to the problem of the transformation of fat into carbohydrate in the organism. M. HENZL. *Z. physiol. Chem.* 195, 248-54 (1931); cf. *C. A.* 25, 494. — In formulating the reactions involved in the transformation of fat into carbohydrate the chief difficulty is the fact that β -oxidation of fatty acids yields 4 C chains such as $\text{AcC}_2\text{H}_5\text{CO}_2\text{H}$ and $\text{MeC}_2\text{H}_4\text{OHCH}_2\text{CO}_2\text{H}$, whereas 3 C chains such as AcCO_2H and AcCHO are required for glucose synthesis. The newly discovered reaction product of $\text{AcC}_2\text{H}_5\text{CO}_2\text{H}$ and AcCHO , viz., diacetyl ethanol, $\text{AcC}_2\text{H}_5\text{C}(\text{Ac})\text{CHOH}$, contains 6 C atoms and by addition of 2 H_2O and 1 O could yield $\text{C}_6\text{H}_{12}\text{O}_6$. Again, this hydroxydiketone could oxidize to 2 AcCHO , thus effecting the transition from the 4 C to the 3-C chain. Feeding experiments with this ketol show that large amts. are tolerated and frequently disappear entirely. Likewise liver pulp removes the ketol. In such expts. the sugar is first destroyed by yeast fermentation which does not attack the ketol. Disappearance of the latter is then shown by the absence of reducing power. Oxidation by KMnO_4 converted the ketol into AcCO_2H , identified by prepn. of the *p*-nitrophenylhydrazone.

A. W. DOX

The fate of choline in the blood. Comment on the paper of Irvine H. Page and Erich Schmidt. FRITZ WARNE AND ERNST BAUSCH. *Z. physiol. Chem.* 195, 255-9 (1931). — The disappearance of added choline, observed by Page and Schmidt (*C. A.* 25, 522) could not be corroborated. Whether fresh blood, serum or blood heated to 55° was used and whether the analysis was performed immediately or 7 hrs. after adding the choline, the recovery as chloroaurate by the authors' method (*C. A.* 23, 5403) was quite uniformly approx. 80%. Positive evidence by isolation and chem. identification of the substance should outweigh neg. results by a bio assay.

A. W. DOX

The occurrence of ether-insoluble lecithins in the brain. WILHELM MERZ. *Z. physiol. Chem.* 196, 19 (1931). — From the protoplasmic fraction of the brain a white hygroscopic substance was isolated which resembled sphingomyelin in its solv. in Et_2O , petr. ether and MeAc . It becomes viscous at 65-72° and melts above 140° with decomposition. Hydrolysis by $\text{Ba}(\text{OH})_2$ in MeOH yielded choline which was identified as the chloroaurate. Choline, the basic constituent of cephalin, was not found. The fatty acids identified were palmitic, very little stearic, and about 21% oleic. Since less than 50% of unsatd. acid was found, the existence of natural lecithins contg. only satd. acids is demonstrated. Of the glycerophosphoric acid present 54% was the β -acid identified by its d.5-soluble salt with $\text{Ba}(\text{NO}_3)_2$. This is the first instance of a naturally occurring lecithin insol. in Et_2O .

A. W. DOX

The follicular hormone. Preliminary paper. BOLESŁAW SKARZYŃSKI. *Z. physiol. Chem.* 196, 19 (1931). By the method of Marnett, in which 5% KOH is used, a hormone was obtained from gravid urine which m. at 246-50° and showed the same ultra-violet absorption curve as Butenandt's hormone. Its activity was 9,500,000 mouse units as detd. by a 50% response after 5 injections during 30 hrs. The substance in KOH soln. showed a complete loss of selective absorption spectrum after 72 hrs., indicating a deep-seated change. Possibly the crystals represent an inert substance to which the hormone is added as a difficultly separable impurity.

A. W. DOX

The ammonia content of frog muscle and the reversibility of ammonia formation in the isolated frog muscle. GUSTAV EMMER. *Z. physiol. Chem.* 196, 23-49 (1931). — A lengthy and tedious polemic against Parnas (*C. A.* 25, 995).

A. W. DOX

The development and phosphatase activity *in vivo* and *in vitro* of the mandibular skeletal tissue of the embryonic fowl. HONOR BRIDGET FELL AND ROBERT ROBISON. *Biochem. J.* 24, 190C (1930) cf. *C. A.* 24, 1672. — By the sixth day of embryonic development the future histological structure of the non-ossifying part of Meckel's rod and of a future cartilage bone such as the palato-quadrato or femur is already detd., although at this stage no histological differences between the 2 types of cartilage have yet appeared. The membrane bone surrounding Meckel's cartilage possesses a high phosphatase activity.

BENJAMIN HARKOW

What effect have the hormones of pregnancy on the growth of the fetus and the changes of pregnancy of the mother? F. SUGERT. *Arch. Gynakol.* 143, 72-9 (1930). — Both the ovarian hormone and the hormone from the anterior lobe of the hypophysis are increased during pregnancy. The ovarian hormone is excreted in the urine in increasing amts. as pregnancy progresses while the hypophyseal hormone is excreted in large amts. from the very onset of pregnancy and in lesser amts. as pregnancy progresses. While in early pregnancy the blood always gives a pos. test for ovarian hormone, at the time of delivery the blood of the mother gave a pos. reaction in 57.7% of the 52 cases studied and the blood of the child a pos. reaction in 76.5%. The content of hypophyseal hormone is the same in the blood of both mother and child and shows no decrease at the time of delivery. It is probable that the ovarian hormone plays an

important role in the changes in the organism of the mother that accompany pregnancy and that it is excreted in larger amts. in the second half of pregnancy because its biologic importance is now lessened. The biologic activity of the hypophyseal hormone during pregnancy must remain the same or increase, as in spite of increased excretion in the urine at the end of pregnancy there is no change in its content in the blood. The hypophyseal hormone is probably directly concerned with the growth of the fetus. The regulation of the amts. of hormones seems to be a function of the placenta, in fact it is most probable that the placenta takes on the function of producing a substance identical with the ovarian hormones and through the fetal portion of the placenta stimulates the anterior lobe of the hypophysis to increased activity. HARRIET F. HOLMES

The calcium content of the blood serum during pregnancy. M. ADLER *Arch Gynäkol* 143, 236-47(1930).—By use of the method of Kramer Tisdall the Ca content of the blood serum was detd. for 14 healthy nonpregnant women, 26 healthy women in different stages of pregnancy, 6 women during labor and 12 women during the puerperium. During pregnancy the Ca content of the serum is lowered and the decrease begins as early as the 3rd month. The decrease is uniform from the 6th to the 9th month. The Ca content rises slightly in the 10th month without reaching the normal level. During labor the Ca content increases to the lower limit of physiologic variability and the degree of increase seems to be related to the duration of the labor pains. Twenty four to 48 hrs. postpartum the Ca content drops below the value before labor. It remains below normal until the 8th day postpartum and then rapidly returns to the lower normal limit. HARRIET F. HOLMES

Potassium and calcium in the menstrual cycle and during pregnancy. R. SPIEGLER *Arch Gynäkol* 143, 218-71(1930).—The K content of the blood undergoes a marked alteration during the menstrual cycle. As a rule, the lowest K value is reached during menstruation and then increases progressively in the postmenstrual period to reach its max. in the premenstrual period. In contrast the Ca content of the blood is very stable and shows no noteworthy change during the menstrual period. The opposite behavior of the K and Ca ions results in a marked change in the K:Ca ratio. In pregnancy the K content of the blood tends to rise in 60% of the cases. The Ca content is not altered with the onset of pregnancy but during the course of pregnancy there seems to be a slight decrease in Ca content. With increasing pregnancy the K:Ca ratio is slightly shifted in favor of the K. In comparing the blood of mother and child the ions of the blood of the child are not considerably increased over those of the mother. The increase affects both K and Ca but the K:Ca ratio in general is increased in favor of K. HARRIET F. HOLMES

The modification of the alkali reserve of the blood by pregnancy, labor and puerperium and its behavior in the newborn. J. MALPATTI AND J. BURTSCHER *Arch Gynäkol* 143, 272-309(1930).—The alkali reserve of the blood decreases during pregnancy. The decrease becomes more marked during labor and is greater in primiparas than in multiparas. The alkali reserve of the blood increases postpartum but does not reach normal value during the first 3 or 4 days. Lactation modifies the CO_2 -combining power of the blood as is shown by a decrease in the alkali reserve of the blood if the blood is withdrawn at least more than 12 hrs. after the beginning of lactation. If the blood is withdrawn before this time, this decrease cannot be demonstrated. After the decrease in alkali reserve of the blood produced by lactation, there occurs a further increase in CO_2 -combining power, which does not return to normal even 12-21 days postpartum. In a case of vomiting of pregnancy and in a case of *nephropathia gravidarum* the alkali reserve was decreased below that of healthy pregnant women. Septic conditions cause a particularly low value for the alkali reserve. The alkali reserve of the blood of the newborn child is in general lower than that of the mother. The newborn child in the first 21 days of life shows a distinct acidosis that is in general greater than that of the mother during the puerperium. HARRIET F. HOLMES

The iodine content of the blood in women under physiologic conditions. W. SCHIERINGER *Arch Gynäkol* 143, 319-37(1930).—The total I content of the blood (inorg. and org. bound I) was detd. in 18 healthy women. A diet poor in I caused a decrease of about 25% in the I content of the blood. After more than 2 days of a diet poor in I, the I content of the blood tends to remain const. with an av. value of 87%. Seasonal variations from winter to summer were not observed. That the menstrual cycle has an effect on the I content of the blood was confirmed. At the onset of menstruation the I content of the blood is increased to 20-25% of the I content in the intermenstrual period. Several normal women were studied more carefully with detn. of basal metabolism and the Reid-Hunt reaction and the I content of the blood was

pura lutea in the ovaries had no stimulating effect on the development of mammalian tissue.

The relation between contraction frequency and lactate accumulation and its bearing upon the economy of tension production and maintenance in striated muscle. M. B. VISSCHER AND P. W. SMITH. *Am J Physiol* 95, 121-9 (1930).—Lactic acid production in the frog's gastrocnemius muscle is inversely proportional to the frequency of stimulation. The lactic acid coeff. of developed tension is a more const. factor than is the coeff. of maintained tension. The slow removal of lactic acid is of advantage in those muscles where maintenance of tension is an important function. Fatigue contracture seems to be due to a depletion to a low level of the store of base for neutralizing free lactic acid.

The source of energy for anaerobic contraction in glycogen-poor muscle. P. W. SMITH AND M. B. VISSCHER. *Am J Physiol* 95, 130-8 (1930).—Lactate accumulation in anaerobic contractions to fatigue is entirely balanced by the loss of glycogen in glycogen-poor, as well as in glycogen-rich muscle. The constancy of the developed tension coeff. of lactate at widely varying levels of glycogen content strongly supports the hypothesis that the reactions associated with lactic acid production yield the energy for the contraction process.

Maternal and fetal blood-sugar changes under various experimental conditions. S. W. BRITTON. *Am J Physiol* 95, 178-83 (1930).—In early pregnancy, blood sugar changes seem to be readily effected in the fetus in response to changes in the maternal blood. In the later stages of gestation, fetal glucose variations are brought about (and then only slightly) only after severe maternal disturbance, such as insulin hypoglycemia, emotional excitement, adrenaline or glucose administration.

Physiological variations of the cardiac output in man. X. The effect of variations in the environmental temperature on the pulse rate, blood pressure, oxygen consumption, arteriovenous oxygen difference and cardiac output of normal individuals. A. GROLLMAN. *Am J Physiol* 95, 263-73 (1930), cf. *C. d.* 24, 5821.—From 0° to 45° the pulse rate rises progressively, blood pressure falls about 10%, metabolism declines to a min. at about 30°, after which it rises slowly, cardiac output is const. between 0° and 28°, above which the output gradually increases. XI. The pulse rate, blood pressure, oxygen consumption, arteriovenous oxygen difference and cardiac output of man during normal nocturnal sleep. *Ibid* 274-84.—Of the functions studied only the pulse rate rises as a result of awakening of the subject. XII. The effect of the menstrual cycle on the cardiac output, pulse rate, blood pressure and oxygen consumption of a normal woman. *Ibid* 96, 1-7 (1931).—No variations in the functions studied were observed as a result of the menstrual cycle. XIII. The effect of mild muscular exercise on the cardiac output. *Ibid* 8-15.—The cardiac output in very mild exercise bears no simple relation to the O₂ consumption.

The mechanism of water exchange in the animal organism. I. The nature and effects of superficial burns. F. P. UNDERHILL, R. KARSNOW AND M. E. FISK. *Am. J. Physiol* 95, 302-14 (1930).—Skin burns, involving about 30% of the total area, were made on anesthetized rabbits by means of a hot iron. The subsequent edema and absorption of the fluid are described. In the production of such a burn heat was shown to penetrate the body sufficient to raise the temp. of the body cavities temporarily to about 111°F. This is regarded as sufficient to lead to local circulatory changes inducing the formation of ulcers, hemorrhages, etc., which have been ascribed to the effects of a burn toxin. II. Changes in capillary permeability induced by a superficial burn. *Ibid* 315-24.—Under the conditions of a burn, capillary permeability is increased. Substances which normally do not pass the capillary wall (e. g., dyes) are found in the edema fluid. Since absorption from the burned area is much slower than normal, increased capillary permeability in one and decreased permeability in the opposite direction may exist. III. The extent of edema fluid formation induced by a superficial burn. F. P. UNDERHILL, M. E. FISK AND R. KARSNOW. *Ibid* 325-9.—With a burn involving about 17% of the total body surface the max. water loss in the edema fluid was about 70% of the total blood vol. IV. The composition of the edema fluid resulting from a superficial burn. F. P. UNDERHILL AND M. E. FISK. *Ibid* 330-3.—Edema fluid, produced by an extensive skin burn, is approx. the same in compn. as the serum of the blood of the burned animal, non-protein N, K, Mg and inorg. P are generally higher in the edema fluid than in normal serum, K may be several times higher. Both the serum of the burned animal and the edema fluid contain less globulin than does the serum of a normal non-fasting animal. V. The relationship of the blood chlorides to the chlorides of the fluid produced by a superficial burn. F. P. UNDERHILL, M. E. FISK AND R. KARSNOW. *Ibid* 334-8.—Blood Cl shows no diminution as long as blood concn. is main-

tained within normal limits. Cl losses may be replaced to a considerable extent. As much as 20% of the NaCl of the blood may be lost to the edema fluid without causing alteration in the Cl content of the blood. VI. The composition of tissues under the influence of a superficial burn. *Ibid* 339-47.—Rabbits were burned under anesthesia, about 17% of the skin area being affected, and allowed to live without food or water for 2-3 days. In comparing the water, ash and chloride content of the tissues of burned animals with those of unburned controls, no effect of the burn on the compn. of the tissues was apparent. Presumably the compn. of the tissues is conserved in preference to that of the blood. When a portion of the skin is burned the Cl content of the entire skin is markedly increased. In muscle tissue injured by a burn the H₂O content is increased, the ash content varies but little from normal, but the Cl content is increased even more than in burned skin. The resorption of material from a burned area is very slow. VII. Dehydration produced by various means. F. P. UNDERHILL AND M. E. FISK. *Ibid* 348-63.—When rabbits are deprived of food and water and in addn. given pilocarpine or hypertonic solns. or bled, the skin may lose 20% of its water, the blood may become concd., and the muscles may lose some water, but a marked loss from all the tissues except the liver (and possibly the kidney), which never lose, results in death. VIII. Dehydration by pilocarpine under varied dietary conditions. *Ibid* 354-70.—The repeated injection of pilocarpine causes pulmonary edema and blood concn., a loss of water from the skin and an increase of water in most of the organs of animals receiving fresh vegetables as part of their diet. J. F. LYMAN

The respiratory quotient of resting mammalian muscle as shown by the eviscerated decapitated cat. L. IAVING AND H. C. FOSTER. *Am J. Physiol* 95, 429-45 (1930).—When over ventilation was avoided and correction made for any change in CO₂ content of the muscles, the R. Q. of the eviscerated decapitated cat was about 0.89. Intravenous injection of glucose caused a slight rise in the R. Q. to 0.85. After large doses of insulin and glucose the corrected R. Q. rose to 0.94, indicating the diversion of oxidation to the principal use of carbohydrate. The action of insulin and glucose, in addn. to increasing carbohydrate oxidation and glycogen storage, affects the CO₂ equil. in muscle, probably increasing muscle capacity for CO₂. J. F. LYMAN

Metabolism. VIII. The effect of estrin injections on the basal metabolism, uterine endometrium, lactation, mating and maternal instincts in the adult dog. M. M. KUNDE, F. E. D'AMOUR, A. J. CARLSON AND R. G. GUSTAVSON. *Am J. Physiol* 95, 630-40 (1930), cf *C. A.* 22, 113.—Basal metabolic rate in normal dogs is not changed by the subcutaneous injection of the female sex hormone, estrin. In normal and castrated bitches injection of the hormone causes estrus and lactation. J. F. LYMAN

The effect of simultaneous injections of the female and male hormones in capons. MARY JUTIN, F. E. D'AMOUR AND E. B. WOMACK. *Am J. Physiol* 95, 641-8 (1930).—The continued daily administration of both hormones together, or of alternate injections, on successive days, resulted in birds with female plumage and male head furnishings. There was no evidence of antagonism between the sex hormones. J. F. L.

The nature of the nerve impulse. I. The effect of carbon monoxide on mediated nerve. F. O. SCHMITT. *Am J. Physiol* 95, 650-61 (1930); cf *C. A.* 25, 732.—Conclusion. The action potential of nerve, whether anaerobic or aerobic, requires an oxidation or oxygenation of substances in nerve, and activation of the O₂ by a respiratory enzyme similar to that of Warburg is essential. J. F. LYMAN

The hormone of the adrenal cortex. F. A. HARTMAN, KATHARINE A. BROWNELL AND W. A. HARTMAN. *Am J. Physiol* 95, 670-80 (1930), cf *C. A.* 25, 1884.—Cortin can be prep'd by extg the adrenal cortex with Et₂O, which dissolves little adrenaline. The Et₂O is dist'd., and the residue ext'd with warm 80% EtOH. Much inert material is removed by chilling and filtering. After removing the alc. the residue (cortin) is dissolved in water. When it is given subcutaneously to adrenalectomized cats, animals, which otherwise would lose wt. and die, carry out all the life functions normally. A case of Addison's disease with a systolic blood pressure of 50 mm. and a pulse of 120 with other characteristic symptoms was revived by the use of cortin. J. F. LYMAN

Changes in sugar and lactic acid content of blood caused by burns. M. A. SLOCUM AND R. D. LIGHTBODY. *Am J. Physiol* 96, 35-9 (1931).—The concn. of sugar and lactic acid in the blood of rabbits that had been subjected to burns with the adrenal glands intact, and after removal of the glands, indicate that increased activity of the adrenals in a burned animal does not account for the observed increase in blood sugar. J. F. LYMAN

The oxygen consumption of isolated muscles for isotonic and isometric twitches. ERNST FISCHER. *Am J. Physiol* 96, 78-88 (1931).—Isotonic twitches with large loads, where only little shortening occurs, have a higher O consumption than isometric

twitches, while isotonic twitches with small loads, and consequently large shortenings, need less O than the isometric twitches. The value of 0.380 for the O consumption per g tension and unit muscle length was obtained, with a max. efficiency of 13.8%. The energy liberated in a twitch depends on 2 factors, the length of the fibers during contraction and the work done. J F LYMAN

The hemoglobin content of the blood of fowls. H H DUKES AND L H SCHWARTZ. *Am J Physiol* 96, 89-93(1931).—By using the Newcomer acid hematin method with the Bausch and Lomb "Improved Newcomer Model" hemoglobinometer, a correction for turbidity disturbance in fowl blood must be applied as follows $C = 0.91 U - 1.49$, where C is the corrected reading and U the uncorrected reading. A corrected value of 9.8 g hemoglobin per 100 cc. of hen blood was obtained, for pullets 8.9%, cocks 13.5, pheasants 13.7, wild ducks 14, wild geese 14.9, brants 14.7, swan 13.4, peafowls 12.0, turkeys 10.7. J F LYMAN

The oxygen pulse in athletic girls during rest and exercise. ELLEN M RADLOFF. *Am J Physiol* 96, 126-31(1931).—From 2 to 3 times as much O_2 is used per systolic discharge of the heart during exercise as during rest. J F LYMAN

The adrenal cortex. L The effect of a lipid fraction upon the life span of adrenalectomized cats. W W SWINGLE AND J J PFIFFNER. *Am J Physiol* 96, 153-63 (1931).—A crude lipid fraction that is capable of prolonging life in an animal whose adrenals have been removed can be prep'd as follows. Fresh adrenal cortex material is ext'd 24-72 hrs at room temp with 2.5 vols of 95% EtOH. The alc is removed through muslin and filtering. The residue is re-ext'd with 2 vols of 80% EtOH. The alc. is dist'd. *in vacuo* to about 7% of the original vol., each ext'n is handled separately. Each concentrate is ext'd 3 or 4 times with an equal vol of benzene for each ext'n. The benzene is dist'd *in vacuo*, the last traces being removed by the add'n of 50-100 cc portions of abs EtOH and evap'g to dryness. The lipid residue is dissolved in corn oil or olive oil with the aid of abs EtOH, the alc being removed by dist'n *in vacuo*. 1 cc represents 30 g fresh cortex. Adrenaline is present in the prep'n. II An aqueous extract of the adrenal cortex which maintains the life of bilaterally adrenalectomized cats. *Ibid* 164-79.—The lipid fraction, as obtained by the method described above, weighing 81 g was treated with 500 cc. acetone in the refrigerator for 24 hrs. The residue was ext'd once with 500 cc. acetone and once with 100 cc. acetone. The acetone was dist'd *in vacuo*, leaving 14.7 g residue. This residue was treated in a separatory funnel with 30 cc petr ether (b 30-60°), 74 cc of 95% EtOH and 26 cc of H_2O . The 70% alc. layer was washed 5 times with 30 cc portions of petr ether. The petr ether soln. and washings were dist'd. *in vacuo* and the distribution procedure was repeated. The petr ether soln., resulting from the second distribution, was ext'd twice with 70% alc. The alc. solns were washed successively 5 times with 30 cc portions of petr ether in the order 2, 3 and 4 distributions, and finally the third and fourth distributions were washed twice with 30 cc portions of petr ether. The active material, sol in 70% alc., weighed 1.489 g. The alc. was removed *in vacuo* to a vol of about 65 cc. and H_2O was added to 100 cc, 1 cc. representing 30 g fresh cortex. The ext was kept in the refrigerator overnight, centrifuged and decanted through a Seitz filter. The clear reddish brown ext contained 0.5809 g of solids, $pH = 5.2$, 1 cc. = 0.36 mg adrenaline, i. e., 99.6% of the adrenaline in the original tissue was eliminated. III. The revival of cats prostrate from adrenal insufficiency with an aqueous extract of cortex. J J PFIFFNER AND W W SWINGLE. *Ibid* 180-40.—Adrenaline was removed from the ext. of adrenal cortex described above as follows. The 70% alc.-sol. fraction contg. 1.49 g solids and 36 mg of adrenaline was dist'd *in vacuo*, the residue dissolved in 100 cc. of 95% EtOH and the soln filtered through two 30 g portions of permutite, at the rate of 1 to 2 drops per sec. The permutite was washed with one 100 cc and one 300 cc portion of 95% alc. The alc. filtrate contained 0.41 g solids and less than 1 mg adrenaline. It was conc'd to about 100 cc and filtered through two 15 cc portions of permutite. The filtrate contained 0.40 g solids and about 0.05 mg adrenaline. The alc. soln was conc'd to about 100 cc and 70 cc water added. The alc. was removed and the ext dild to 100 cc with water. The resulting milky soln was clarified and sterilized by filtering through a Seitz filter. The ext. was crystal clear and very pale yellow. It contained 0.29 g solids, 0.05 mg adrenaline, $pH = 5.65$. One cc. = 30 g fresh beef adrenal cortex. Large doses of this ext can be given without any deleterious effects. Cats in collapse from adrenal insufficiency, following removal of the glands, were restored to a normal condition after injection of the ext. J F L.

Water diuresis. D McK. RICH. *J Physiol* 70, 45-52(1930).—Large vols. of fluids were given through the stomach to normal dogs and the effects on blood and urine observed. There was (1) a const lag of diuresis behind the changes in the blood

following the ingestion of isotonic salt solns; (2) a constancy of changes in the blood following the ingestion of isotonic salt solns and (3) an inconstancy of changes in the blood following the ingestion of water. Dtn of the electrolyte concn of the blood is probably the responsible factor in instigating a water diuresis. J. F. LYMAN

Water diuresis. IV. Changes in the concentration of electrolytes and colloids in the plasma of decerebrate dogs produced by the ingestion of water. L. E. BAYLISS AND A. R. FRANK. *J. Physiol.* 70, 163-6 (1930), cf. *C. A.* 24, 4821.—Conclusion: In dogs the absorption of water from the digestive system results in a fall in the concn of plasma colloids, which is approx. that which would take place if the extra water were uniformly distributed throughout the tissues. There may be a fall in the electrolyte concn of the plasma, but this is always less than that of the colloid concn. Intravenous injection of 0.6% NaCl soln also results in a fall in the concn of the plasma colloids of the expected magnitude. J. F. LYMAN

The respiratory quotient of the eviscerated spinal cat. A. B. CORKHILL, H. H. DALE AND H. P. MARKS. *J. Physiol.* 70, 85-97 (1930).—The R. Q. of the decapitated cat, whose viscera have been removed and whose blood sugar is maintained by artificial addn of glucose, is usually found to be 1. Somewhat lower values are sometimes found. Excessive ventilation in these expts was excluded. J. F. LYMAN

Delayed anaerobic heat production of stimulated muscle. H. BLASCHKO. *J. Physiol.* 70, 105-109 (1930).—The occurrence of delayed anaerobic heat after a tetanic stimulus in muscle has been verified. It seems to be due to some real process going on in the muscle, but condensation of water vapor, caused by a temporary rise of osmotic pressure, and also the non uniform heat production in contraction or control have been excluded as possible causes. J. F. LYMAN

Reactions of denervated voluntary muscle, and their bearing on the mode of action of parasympathetic and related nerves. H. H. DALE AND J. H. GADSDEN. *J. Physiol.* 70, 109-44 (1930).—Evidence is obtained, from a study of the action of drugs on denervated muscle, that vasodilator effects of impulses in the parasympathetic and dorsal root nerve fibers depend upon peripheral liberation of acetylcholine. J. F. LYMAN

The nature of inhibition in the intestine. D. FINKLEMAN. *J. Physiol.* 70, 145-57 (1930).—When the movements of a certain piece of intestine were inhibited by nervous stimulation, a substance appeared in the liquid passing over the surface of the intestine which had the power of inhibiting a 2nd piece of intestine. Conclusion: The inhibitory nerves to plain muscle act by liberating an inhibitory substance. The effects of this substance are, in all ways, similar to those produced by adrenaline. J. F. LYMAN

The automatic regulation of gastric acidity. F. L. APPERLY AND JOAN H. NORRIS. *J. Physiol.* 70, 155-68 (1930).—Changes in acidity of acid solns placed in the stomach were studied. Conclusion:—The reduction in acidity observed with solns above a certain concn takes place mostly by diln, the diluting fluid being partly secreted by the gastric mucosa and partly regurgitated from the duodenum, and, in the majority of cases, the reduction in acidity is partly effected by duodenal alkali in addn. J. F. LYMAN

Interrelations of respiratory and gastric secretion. C. E. BAUNTON AND M. C. G. JARRELS. *J. Physiol.* 70, 184-94 (1930).—If gastric HCl is derived from the blood chloride thus: $\text{BCl} + \text{H}_2\text{CO}_3 \rightarrow \text{BHCO}_3 + \text{HCl}$, an increased CO_2 tension in the alveolar air might be expected during gastric secretion with a subsequent fall during the phase of pancreatic digestion. No such const. shifts were observed. It is argued that any possible excess of alkali resulting from the secretion of amts. of HCl normally produced might be dealt with without necessarily causing a measurable rise in alveolar CO_2 . J. F. LYMAN

Lactic acid formation and removal with change of blood reaction. M. GRACE EGGLETON AND C. L. EVANS. *J. Physiol.* 70, 261-8 (1930).—The lactic acid content of the blood, used to perfuse isolated mammalian muscle, is increased by increasing the alk. of the perfusing fluid. When the acidity of the fluid is increased, the lactic acid falls again, if the blood is perfused through muscle, but not if perfused through lung alone. J. F. LYMAN

The lactic acid content of the blood after muscular contraction under experimental conditions. M. GRACE EGGLETON AND C. L. EVANS. *J. Physiol.* 70, 269-93 (1930).—The following points were settled: (1) The av. lactic acid content of superficial venous blood is closely parallel to that of the arterial blood, yet individual samples show variations. (2) The lactic acid content of tissues is somewhat less than that of the blood plasma. This difference, up to about 10%, can be accounted for on the basis of the higher water content of the plasma. (3) Lactic acid in the blood and tissues does not come to equl. until at least 10 min. after the end of exercise. (4) The liver removes

lactic acid from the blood, but it is not solely responsible for lactic acid removal (5) In the eviscerated animal (a) lactic acid does not rise during rest during the 4 hrs following the operation, and (b) glycogen synthesis occurs to a considerable extent in muscles during recovery from exercise (6) In the resting condition the sugar and lactic acid content of the blood are normal under amytal anesthesia, but abnormally high in the decerebrate animal

J F LYMAN

The diffusion of creatine and urea through muscle. P. EGGLETON. *J. Physiol.* 70, 294-300(1930)—In salt solns. contg creatine, in which resting muscles are immersed, there is creatine equil. without loss or gain of creatine to the muscle when the liquid contains 80 mg creatine per 100 cc of soln or 65 mg per 100 g of muscle (80 mg per 100 cc of H_2O in the muscle). This indicates that about 20% of the total creatine of resting muscle is in the free state. With fatigued muscle equil. is obtained at concns of 200 to 300 mg creatine per 100 cc of soln. Urea distributes itself between resting muscles and surrounding isotonic salt solns as if the whole of the water in the muscles is capable of dissolving urea.

J F LYMAN

The effect of carbon dioxide on the circulation. I. R. J. S. McDOWALL. *J. Physiol.* 70, 301-15(1930)—Over ventilation has 2 effects: (1) a dilatation of blood vessels in certain regions of the circulation resulting from a loss of tone in the vasomotor center; (2) a constriction in other regions produced by a different mechanism. Accordingly blood pressure may fall, rise or remain unchanged as a result of over-ventilation.

J F LYMAN

Excitable substances in the nerve-muscle complex. W. A. H. RUSHTON. *J. Physiol.* 70, 317-37(1930)—Conclusion: There are 2 different excitable substances in the nerve-muscle complex. The γ substance, which is isochronous with nerve, so far from being the excitable substance of the normal muscle fiber lies in a direction which is usually quite different, and appears to be absent from the pelvic extremity of the sartorius. The α substance, on the contrary, seems to have the properties of the excitable element of the muscle fiber.

J F LYMAN

The carriage of carbon dioxide by blood. M. N. J. DIRKEN AND H. W. MOOK. *J. Physiol.* 70, 373-84(1930), cf *C. A.* 24, 3157.—An app. is described with which it is possible to measure either the p_n or the CO_2 tension of a fluid after CO_2 has been added, at various moments up to about 4 sec. There is evidence that hydration of CO_2 is slow; dehydration of H_2CO_3 , while slow, is much more rapid than its reverse. If hemoglobin is present, anhydrous CO_2 disappears from the fluid just as rapidly as any acid. There must be (1) some direct combination between hemoglobin and anhydrous CO_2 or (2) a catalytic acceleration of the hydration of CO_2 by hemoglobin.

J F LYMAN

Selective absorption of carbohydrates. J. J. R. MACLEON, H. E. MACGEE AND C. B. PURVES. *J. Physiol.* 70, 404-16(1930)—The surviving intestine shows selective absorption for sugars at body temps., but at 4° the intestinal wall behaves like a dead membrane. Selective diffusion could not be demonstrated when mixts. of sugars were used in place of sep. solns. of each sugar. The selective absorption of glucose is most rapid when the concn. of glucose is about 0.75 M.

J F LYMAN

The vapor pressure of normal human blood. R. MARGARIA. *J. Physiol.* 70, 417-33(1930)—Measurements were made by the use of Hill's thermoelectrical method (*C. A.* 24, 3695). Expressed in terms of g. of NaCl in 100 g. of H_2O , the values were for men (19 observations) 0.9447 ± 0.00495 , for women (16 observations) 0.9269 ± 0.0059 . The reality of the difference between the sexes is statistically certain. The difference is due mostly to bicarbonate and urea differences. After drinking 1500 to 2000 cc. of water, values as low as 0.88 were noted and after severe muscular exercise a value of 1.048 was observed.

J F LYMAN

The absorption of calcium from the gall bladder. L. M. G. SACHS AND T. SCIRIRE. *J. Physiol.* 70, 434-40(1930)—Ca can be absorbed from the gall bladder, probably along with water and other substances. It is suggested that there is a "Ca circle" in the liver, which provides a mechanism for a continuous supply of Ca to the general circulation.

J F LYMAN

The nervous control of insulin secretion. J. HOET AND H. ERNOULD. *Proc. Physiol. Soc., J. Physiol.* 70, i, ii(1930)—A demonstration that vagus innervation of the pancreas is necessary for normal insulin secretion is described. When 60 cg glucose per kg. rabbit is injected into an animal under allylsopropyl barbiturate of diethyl amine whose kidneys have been removed the blood sugar level comes back to 0.130 or 0.140%, but when also the vagi are cut, the glucemia will stay for hrs. at a level of 0.300 or 0.400%.

J F LYMAN

Glycogen recovery in mammalian muscle as an insulin function. G. DEBOIS. *Proc. Physiol. Soc., J. Physiol.* 70, ii-iv(1930), cf *C. A.* 25, 1684.—The glycogen content

of muscles from cats, whose metabolism had been altered by operations, was detd. immediately after periods of work and after a recovery period of 1.5 hrs. Conclusion: In the absence of liver and pancreas and other visceral organs there is no recovery of muscle glycogen after loss by exercise. There is no recovery in a cat under ether with the vagi cut. With the vagi intact there is a normal recovery. After removal of the pancreas there is no recovery whatever. Insulin injection makes recovery possible in a very short time. Glycogen recovery after contraction in muscle is an insulin function. J. F. LYMAN

A possible nervous mechanism involved in the liberation of histamine. R. G. MACCARSON AND S. PEAT. *J. Physiol.* 71, 31-5 (1931).—Cats were anesthetized, the left lung was removed for analysis, the animal then subjected to the exptl. conditions, after which the right lung was removed and assayed. The method used was capable of detecting a difference of 20% or more. Neither stimulation of the nerves supplying the lungs, nor perfusion with adrenaline, pilocarpine or physostigmine gave rise to any detectable change in their histamine content. J. F. LYMAN

Ovulation. IV. Induction of ovulation in the hypophysectomized rabbit by administration of anterior lobe extracts. MARGARET HILL AND A. S. PARKES. *J. Physiol.* 71, 36-9 (1931), cf. *C. A.* 24, 5819, 25, 2765.—A hormone, extd. from the anterior lobe of ox. pituitaries, replaces the function of the removed pituitary in the rabbit in causing ovulation in estrous, pseudo-pregnant or immature rabbits. V. The action of the ovulation-producing substance of urine of pregnancy on the hypophysectomized rabbit. *Ibid.* 40-6.—1 pts. of urine from pregnant women induced ovulation in 4 out of 19 trials in hypophysectomized rabbits. 1 pts. of normal rabbits' placentas were inactive, but highly active exts. were obtained from the placentas of intact or hypophysectomized rabbits previously injected with exts. of pregnant urine. J. F. L.

Adaptations of the organism to changes in oxygen pressure. L. PHYSICO-CHEMICAL changes in human blood at low oxygen pressure. D. B. DILL, H. T. EDWARDS, A. FÖLLING, S. A. OSERO, A. M. PAPPENHEIMER, JR. AND J. H. TALBOTT. *J. Physiol.* 71, 47-63 (1931).—Measurements were made at sea level, 10,000 ft. and 14,000 ft. The O transport capacity may be decreased by 20% in normal men at 10,000 ft. altitude. This large effect may be due to an indirect effect of a small change in O satn. of arterial blood upon the limiting value of cardiac output. J. F. LYMAN

The initial and recovery heat production of vertebrate nerve. D. W. BROWN. *J. Physiol.* 71, 136-44 (1931).—Heat production of frog nerve has been measured with improved app. The recovery heat production is at a max. rate immediately after a stimulus of 9-15 sec. and continues for some min. (11 min. at 19° and 7.75 min. at 24-5°) at a decreasing rate. The initial heat production is 8.9% of the total. J. F. L.

The creatine and phosphorus contents of muscle. MARION BROWN AND C. G. LEMERIE. *J. Physiol.* 71, 214-21 (1931).—When the concn. of creatine in the muscles of cats was increased, following the absorption of creatine from the intestine, the concn. of the total acid sol. P was also increased. J. F. LYMAN

Studies on the action of the pH in striated muscle and on their buffering power. S. GOLDBRECHT. *Arch. sci. biol. (Italy)* 15, 505-24 (1930).—Varying the H ion concn. of the circulating liquid, whether it be toward the acid or toward the alk. side, increases the threshold of the excitability and the intensity of the stimulus necessary to produce max. contraction, lowers the height of the contraction and leaves unchanged the latent time and the duration of the contraction. The muscle retains its buffering power for over 24 hrs. when using solns. between pH 3.2 and 10.8, while the more acid or alk. fluids cause the loss of such capacity within 45 min. The acid solns. also bring about an inexcitability of the muscle, while the alk. solns. do not. The loss of excitability is related to the transformation of the protein anion into cation. The size of the edema formed is independent of the H ion concn. The variation of the H ion concn. produces a marked vasoconstriction. PETER MASUCCI

Respiratory exchange in asphyxia. GIOACINO MANNERY. *Arch. sci. biol. (Italy)* 15, 523-54 (1930).—The scope of these studies was to clarify certain points in the behavior of the respiratory exchange in man under conditions of asphyxia caused by an abnormal compn. of the respired air. Voluntary apnea, after normal respiration, may continue until the partial pressure of CO_2 and O in the alveolar air reaches a definite value: 51 mm. Hg for CO_2 and 78 mm. Hg for O. Respiration, for 90 sec., of a gaseous mixt. (7% CO_2 and 11% air) similar in compn. to that of alveolar air, at the end of the voluntary apnea causes variations in the compn. of the expired air, pulmonary ventilation, R. Q., elimination of CO_2 and consumption of O, these changes, which are enumerated in detail, take place during and after the respiration of the gaseous mixt. The phenomena are observed in the same manner, but less conspicuously, when a gaseous mixt.

contg 7% CO_2 and 20% O is respired. In this case, during the period of asphyxia there is also a diminution in the elimination of CO_2 , with a successive increase in elimination in the post asphyxia period. For respiration of a mixt contg 11% O without the addn of CO_2 , the curve of O absorption is the same, but with fewer oscillations than in the preceding case. During the period of asphyxia, the mixt has no action on CO_2 elimination, but does cause a slight increase in the post asphyxia period. In the respiration of a mixt poor in O as well as in the respiration of a mixt rich in CO_2 , there is an increase in pulmonary ventilation, but it is less intense than in the respiration of the asphyxia mixt, the increase is greater during the post asphyxia period than during the asphyxia period. In both cases at the end of the 12th min, there is yet an "O debt". The decreased absorption of O in the presence of an excess of CO_2 is probably attributable to a displacement in the curve of the dissociation of oxyhemoglobin as a result of the increase of the CO_2 tension in the blood. PETER MASUCCI

Variations in certain fermentative properties of fatigued muscles. EMILIO BECCARI. *Boll. soc. ital. biol. sper.* 5, 1180-4 (1930).—The aim was to see whether any quant. differences existed between the glycolytic and amylolytic properties of fatigued muscles and normal muscles. Fatigue was induced by periodic elec. current stimulation of white mice. The animals were then bled to death, the dorsal muscles of the legs removed and triturated to a pulp. This pulp was suspended in physiol. saline at pH 7.6 and placed in the incubator at 37°. The substances fermented were glucose and rice starch. The glycolytic power of fatigued muscles in contrast to that of normal muscles was inhibited, but the amylolytic power remained unchanged. Since these changes took place even when the lactic acid was removed and when the reaction of the fatigued muscles was the same as the normal, the fermentative action is attributed to the so-called fatigue poisons. PETER MASUCCI

Studies on pressor changes and on certain blood constants under the influence of mud baths applied in various ways. C. CIPRIANI AND A. ROBECCHI. *Minerva med.* 1931, 1, 269-75.—The diffused vascular action of mud baths is accompanied after each application by marked changes in the general arterial pressure, consisting in a lowering of the pressure followed by a return to the original values. In hypertension cases, the lowering is more marked and may lead to an advantageous decrease of pressure at the end of the treatment. The variations in arterial pressure are simultaneous to general changes in the blood, during the first phase there are transitory signs of blood serum dilution, and later this is followed by an increased density of the serum. P. M.

Determinations of blood volume in human beings. III. Preliminary investigations. IV. Results with the vital red method. C. C. FLEISCHER-HANSEN. *Skand. Arch. Physiol.* 59, 243-63 (1930).—Injections of 0.5 mg. vital red per kg. of body wts. are used under conditions which insure even distribution of the dye in the blood stream. The mean blood wt. in 18 men was found to be 5.1% of the body wt., while in 12 women the results varied from 4.8 to 10.5%, the high results being apparently due to elimination of the dye. S. MORGULIS

The fat of sow milk. OTAKAR LAXA. *Ann. fols.* 24, 87-8 (1931).—At ordinary temp. the fat consists of a granular mass similar to melted butter, light yellowish brown, with a pig like odor, and having the following consts.: solidifying point 17-18.5°, m. p. 28°, n_D^{20} 1.52 (presumably oleo-refractometer degrees), sapon. no. 193.0, Hehner no. 93.7, I no. 58.2, Reichert Meissl no. 2.1, Wauters Polenske no. 1.2. The mol. wts. of the volatile acids indicate that caprylic and capric acids are present in least amts. The insol. acids solidify at 5-37.5°, m. 39-40°, have a mol. wt. of 276.9 and I no. 61.9%. The said acids m. 60.5° and have a mol. wt. of 256. Detn. of stearic acid via Hehner and Mitchell (*Analyst* 1896, 316) gave neg. results. Conclusion: The compn. of the fat is approx. oleic acid 64.5, palmitic acid 26, myristic acid 2.6, volatile acids (caprylic and capric) 1.4, glycerol 4.9%. A. PAPINEAU COUTURE

The unsaponifiable portion of the bile lipoids. E. P. HAUSSLER AND E. BRACCHILI. *Hdt. Chim. Acta* 13, 908-15 (1930).—In an investigation of the estrus hormone, 3 compds. (I, II and III) were isolated from ox bile by repeated extr. and crystns. They were all colorless, insol. in water and KOH, contained neither N nor S, were not pptd. by digitonin, dissolved in CHCl_3 did not add Br_2 and were not d-compd. by KMnO_4 in acid. They differed from cholesterol (IV) in their reaction to the Liebermann-Burchard and the Salkowski tests and they gave neg. results in the Rosenheim, Tortelli, Jaffe, Carr and Price, and Fetteskofer tests. None of the 3 acted as an estrogen when injected into castrated rats. I ($\text{C}_{27}\text{H}_{46}\text{O}_4$ or $\text{C}_{27}\text{H}_{44}\text{O}_4$) occurs as an addn. compd. (V) m. 172-3°, $[\alpha]_D^{25}$ (4.7% in CHCl_3) -32.5°, with IV. The IV was pptd. with digitonin and I, m. 185-7°, was obtained from the mother liquor. Acetylation of V gave no

identifiable product but benzoylation in $\text{C}_6\text{H}_5\text{N}$ gave a resinous substance, sol with difficulty in 1 (OH); it yielded 1, m 191.5°. This 1 gave an Ac deriv m 169-70°. II crystd out during the crystn of 1. It m 217-8° and reacts with Ac_2O but the reaction product could not be isolated. III was obtained at the same time as an amorphous powder which yielded crystals m 225-7°. III appears to have 2OH groups and its Ac deriv m 231-2°, $[\alpha]_D^{25}$ (1% in PhH) -51°. C. H. PEET

Factors influencing the respiration of erythrocytes. I. Primitive avian erythrocytes. G. PAYLING WRIGHT. *J. Gen. Physiol.* 14, 179-89(1930).—The O consumption of normal and "primitive red cells" of fowls' blood was detd during the course of anemia produced by injection of PhNH_2NH_2 . "Primitive red cells" have an O consumption at least 20-25 times greater than normal red cells. Suspension of the cells in NaCl solns of various concns has little effect on the O consumption of the cells. The red cells from anemic blood are sensitive to changes in p_{H} . The max O consumption in NaCl soln occurred at p_{H} 7.75. The red cells were more sensitive to variation of p_{H} on the acid side than on the alk. side. The addn of glucose to the medium increased O consumption of the cells, 0.6% glucose caused a 14% increase above solns which did not contain glucose. Low concns of amino acids were practically without effect on O consumption, higher concns of some of them diminished it. II. Mammalian reticulocytes. *Ibid.* 201-13.—The O consumption of rabbit reticulocytes was detd during anemia produced by PhNH_2NH_2 . Respiration increased greatly during regeneration, but the O consumption per billion cells throughout the period remained approx the same. Respiration of the reticulocytes was affected by changes in p_{H} of the medium in which they were suspended, reaching a max at p_{H} 8, the intracorpuseular p_{H} probably being about 7.75. Variations in toxicity of the suspending medium had little effect on respiration. Glycine, alanine and glucose in the suspending medium produced no acceleration in respiration of the cells. Higher concns of glucose tended to depress respiration. The material oxidized is largely or entirely contained in the cells when they are liberated from the marrow. The O consumption in cu mm per 100 cu mm cells for the fowl and rabbit is as follows: primitive red cell, fowl, 140; normal red cell, fowl, 12; reticulocyte rabbit, 70. C. H. RICHARDSON

G—PATHOLOGY

H. GIDEON WELLS

Speed of reaction of antitoxins and its significance on the curative value of sera. R. KRAUS. *Annales guim farm (Chile)* 1, 13-8(1931).—Review of the controversy between Ehrlich and Kraus and co-workers on the correlation between the antitoxic value of sera as measured *in vitro* and their actual curative value. The Ehrlich method has only a quant. value but does not take into account the quality of the antitoxins. The curative value is intimately connected with the so-called speed of reaction.

C. ABELEDO

Urea administration in water intoxication. V. J. HARDING AND L. J. HARRIS. *Trans. Roy. Soc. Can.* 24, Sect. V, 101-10(1930).—Retention of water to produce convulsive symptoms in dogs must exceed 60 cc. per kg. body wt. All animals in such convulsions recover by administration of 10% NaCl, but only 2 of 14 recovered by administration of hypertonic urea solns. It is doubtful if the convulsions of epilepsy or eclampsia are manifestations of water intoxication, even though they may be accompanied by a pos. water balance.

A. T. CAMERON

Effect of certain oxidation-reduction potential indicators on diphtheria toxin. P. J. MOLONNEY AND EDITH M. TAYLOR. *Trans. Roy. Soc. Can.* 24, Sect. V, 127-32(1930).—2-Chloroindophenol, phenolindophenol and similar indicators will render the toxin atoxic; oxidation seems to be a factor in the detoxication. Its rate increases with increased concn. of indicator, and, in the range p_{H} 6.5-7.5 with increased alk. of soln. Toxoids may be produced by such indicators which are equal antigenically to those produced by the action of formalin on toxin.

A. T. CAMERON

Studies on hay fever. R. BRINAGM AND G. FALCONI. *Ann. chim. applicata* 20, 547-58(1930).—The active principle in *Vicia faba* from Sardinia has been shown to be a glucoside having no cyanogen groupings. A phytosterol (vicosterol) was also isolated, as well as sugar whose compn. was not detd. The necessity for studying the pharmacology of the above glucoside and its relation to the pathology of hay fever is indicated.

A. W. CONTIERI

The partition of amino acids in the blood in anemias. ALEXANDER SIMON. *Arch. exp. Path. Pharmacol.* 154, 239-46(1930).—The amino-acid contents of the whole blood and of the plasma in various types of anemias are within normal limits. The amino

acid content of the cells is somewhat elevated, as is the ratio of amino N of the cell to that of the plasma. In the regeneration phase, this ratio returns to its normal values.

II EAGLE

The relationship between creatinuria and muscle glycogen. CARLO BRENTANO. *Arch expil Path Pharmacol* 155, 21-45(1930)—In spontaneous creatinuria, as well as in creatinuria induced by acidosis or narcosis, there is a depletion of glycogen from the skeletal muscles. Agents which cause such a glycogen depletion (adrenaline, phlorhizin, convulsions, CO) cause creatinuria.

II EAGLE

Studies in edema. II. Diseases with change in the colloid-osmotic pressure. HANS HOFSTERS. *Arch expil Path Pharmacol* 155, 248-56(1930). cf C A 22, 1400—The av. value of the colloid-osmotic pressure is about 371 mm H₂O, 40-45 mm. per g. serum protein. In essential hypertension it averages 288 mm., 22-30 mm. per g. protein, the serum protein being usually high. In acute nephritis, the colloid-osmotic pressure averages 250 mm. H₂O, with a low serum protein. Lowest values per g. protein are observed in uremia (23-29 mm. H₂O per g.). In chronic nephritis the serum protein averages 7.7%, the colloid-osmotic pressure, 311 mm. H₂O. In both amyloidosis and cardiovascular insufficiency there is a lowered colloid-osmotic pressure. Normal values are observed in diabetes mellitus except for cases in coma, in which it is very low.

II EAGLE

Level of iodine metabolism, unsanitary conditions of life and goiter. R. MCCARRISON AND CLIVE NEWCOMB. *Indian J. Med Research* 17, 1061-1100(1930)—Young rats on a diet of oatmeal, patent flour, linseed meal, Ca₃(PO₄)₂ and NaCl, contg. 125-250 γ I₂ per kg., did not develop hypertrophic goiter unless they were kept under unhygienic conditions. This type of goiter could be prevented by adding sufficient I₂ to the diet to change the urine I₂ from 33 γ to 85 γ per l. Another type of goiter (lymphadenoid) may develop even on a diet contg. I₂ and under conditions of perfect sanitation. A statistical analysis of the results obtained with the individual rats indicate that these conclusions were justified, also, that unhygienic conditions may cause splenic enlargement in the animals, not affected by I₂. Rats under unhygienic conditions without iodine have comparatively large livers.

II EAGLE

The iodine metabolism in Basedow's disease and the explanation of the post-operative reaction following thyroidectomy. ARTHUR BIER. *Klin Wochschr* 9, 819-21(1930)—Following thyroidectomy, there is a 5-10% fall in blood I, i. e., the post-operative reaction is not due to the mech. expression of the thyroid. Instead, it is considered due to a sudden fall of the thyroid secretion, a "hypothyrotoxic" shock, in an organism on a high level of I₂ metabolism, it is best avoided by continued treatment with I₂ up to and following thyroidectomy. B. points to the fact that the post-operative urine contains only traces of I₂, indicating I₂ insufficiency and retention rather than over dosage.

II EAGLE

The cause of icterus neonatorum. KARL J. ANSELMINO AND FREDRICH HOFFMANN. *Klin Wochschr* 10, 97-100(1931)—A review.

II EAGLE

The blood cholesterol in arterial hypertension. C. ALVAREZ AND S. M. NEUSCHLOSS. *Klin Wochschr* 10, 244-7(1931)—The blood serum of individuals with normal arterial blood pressure is not said with respect to cholesterol (57-90% of the total possible soly.), but in 21 of 25 hypertensives the blood serum was supersatd. (106-132%). The possible etiological connection between hypercholesterolemia and arteriosclerosis is discussed.

II EAGLE

A new form of diabetes mellitus in animal experimentation. M. MAJUS AND O. STERNBERG. *Klin Wochschr* 10, 264-5(1931)—Removal of an adrenal gland from a dog, combined with the surgical implantation of more parathyroid glands, causes the animal to develop glucosuria, and a blood sugar tolerance curve characteristic of diabetes.

II EAGLE

Demonstration of specific antibodies in vitro in severe allergy to fish and yeast. KARTE JAFFÉ. *Klin Wochschr* 10, 304-6(1931)—The sera of subjects sensitive to fish and yeast give complement fixation with an appropriate fresh ext. The sera must be used in the active state, possibly because the antibody is thermolabile. Only patients with a pos. immediate skin test will give the reaction.

II EAGLE

The occurrence of heavy metals in human gallstones. R. SCHÖNHERRER AND W. HERKEL. *Klin Wochschr* 10, 315-6(1931)—Cu, Zn, Mn and Fe are all present in gallstones.

II EAGLE

The clinical significance of the hyperglycemic principle in the pancreas (the so-called initial insulin-hyperglucemia). MAX BORGER. *Klin Wochschr* 10, 351-4(1931).

II EAGLE

The relationship between cancer and the lipid metabolism. II. F. BURGHARDT

AND W. JORL. *Klin Wochschr* 10, 397-8(1931), cf C. A. 23, 3735 —The presence of large quantities of cholesterol in the immediate vicinity of malignant, as contrasted with benign, tumors may be an important factor in detg their further spread. Malignant tumors contain much more cholesterol than normal tissue or benign tumors.

II EAGLE

Phenolsulfonephthalein test in surgical kidney disease, with particular reference to its use in ureteral catheterization. HAYAMI HIROSE. *Tohoku J Exptl Med* 15, 369-97(1930) —The prognostic value of the phenolsulfonephthalein test is stressed as well as the value of ureteral catheterization for the test of a single kidney function.

II EAGLE

The circumvention of the hemolytic and anticomplementary properties of the antigens used in the Wassermann reaction. F. OKOLOV AND L. GRIBANOV. *Z. Immunittts* 68, 7-13(1930) —By filtering the aq. diln of the antigen through a paper filter, its hemolytic and anticomplementary components can be largely removed without affecting its sensitivity for the Wassermann reaction.

II EAGLE

The nature of "H" and "O" agglutinogens. M. N. FISCHER AND R. B. HOCHEBARG. *Z. Immunittts* 68, 43-55(1930) —Agar cultures of *B. typhosus* and *proteus* X 14 were extd with water. These exts combined both "H" and "O" agglutinogens, effecting a coarse- and fine-flaked agglutination, resp. The "H" substance ppts from the ext at its isoelec. point as a single voluminous floccule, while the "O" substance ppts as a fine granular sediment at a somewhat more acid reaction. Both are protein (globulin), contg. the same proportions of C, H, N, S and P, they are considered by F. and H. to be the same substance in a different state of aggregation. The sera of patients with typhoid vary in their pptg. action upon these 2 substances.

II EAGLE

The purification of hemolytic antibodies. HANS V. FILER AND EDVARD BRUNILA. *Z. Immunittts* 68, 121-36(1930), cf C. A. 24, 2179 —By absorbing ambocceptor onto the stroma of red cells, and subsequently extg the agglutinated stroma in weakly alk. reaction (pn 10-11), one obtains highly active solns. By dialyzing these exts. one obtains a 300-fold purification as compared with original serum.

II EAGLE

Disturbances in the equilibrium of the serum proteins in the blood serum of lepers. K. SCHLOSSMANN. *Z. Immunittts* 68, 154-68(1930) —The total protein content of leper serum increases in the early stages of the disease after the appearance of skin manifestations and decreases in the late stages. The serum globulin is increased in fever and during active phases. The lowest albumin/globulin ratios are obtained in *lepra tuberosa* and *mixta*. Those sera giving a positive Wassermann reaction contain large quantities of globulin, but not all sera with increased globulin are Wassermann-pos. In contrast, all syphilitic sera with a low albumin/globulin ratio are Wassermann-pos. The globulin of leper serum is much more stable in the presence of beef heart lipid than that of syphilitic serum.

II EAGLE

The effect of phenol upon specific and non-specific complement-fixation phenomena. KARL DREYFUSS. *Z. Immunittts* 68, 193-210(1930) —The addn of phenol to the cholesterolized LtOil ext. of beef heart as used in the Wassermann reaction causes an increase in its sensitivity. Phenol also makes lecithin, bacterial exts. and bacterial suspensions give complement fixation in the Wassermann reaction, but non-specific positive reactions are thereby obtained. Suspensions of macerated guinea pig heart or aq. exts. acquire Wassermann reactivity by the addn of phenol, without attaining the sensitivity of the alc. exts. Here again, the phenol predisposes to non-specific reactions.

II EAGLE

The effect of intracutaneous stimulation upon agglutinin formation. JENŐ SZÉR. *Z. Immunittts* 68, 274-6(1930) —The intracutaneous injection of a suspension of paratyphoid *B. bacilli* into rabbits is twice as effective in the formation of agglutinins as the subcutaneous injection.

II EAGLE

The blood-group titer in tuberculosis. HERMANN ZANTROP. *Z. Immunittts* 68, 277-85(1930) —There is no significant difference in the isoagglutinin titer of tuberculosis, as compared with normal sera. No significant change in titer is observed during a 3 month rest-cure, aside from normal variations. Anti-A agglutinins are usually stronger than anti-B, and sera of group O with a high anti-A agglutinin content usually have a high anti-B titer also.

II EAGLE

The titration of small quantities of tetanus toxin. B. FRIERABEND. *Z. Immunittts* 68, 286-98(1930) —The indicator used is the local tetanus produced by the intramuscular injection of a mixt. of the serum and the minimal effective amt. of toxin into the hind leg of a guinea pig. As little as 1/500 unit can be detected in this manner.

II EAGLE

Potassium and calcium in anaphylaxis under blockade conditions. H. M. A.

KUSHNAREV *Z. Immunitäts* 68, 299-303(1930); cf. *C. A.* 25, 1898.—Blockade of the reticulo endothelial system of a sensitized guinea pig by the injection of India Ink, carmine or trypan blue 30 min before the re-injection of the sensitizing material, prevents anaphylaxis. The usual prophylactic effect of Ca^{++} in anaphylaxis and the accentuating action of K^{+} are modified when the reticulo-endothelial system has been blocked in this manner. II FAGLE

Amines as antigens. Z V YERMOLYEVA AND I S BUYANOVSKII *Z. Immunitäts* 68, 312-6(1930).—A series of amines injected into rabbits caused the formation of pptg and complement-fixing antibodies which reacted more or less specifically with the amine used for injection ($\text{N}(\text{Me})_3$, heptylamine, heptadecylamine, etc). II FAGLE

Residual antigens of vibrios. Z V YERMOLYEVA AND I S BUYANOVSKII *Z. Immunitäts* 68, 310-50(1930). Twenty-four hr agar cultures of cholera and clear and cloudy vibrios were suspended in saline. Kott was added up to 5%, and the mixt kept overnight in the ice box. Acetic acid was added at boiling temp., and the ppt removed after cooling. The acetic acid pptn was repeated 5 times, the filtrate neutralized, and pptd with 5 vols of 75% EtOH. The EtOH pptn was repeated 8 times in both acid and alk reaction. The dry powder obtained is protein free, reduces Fehling's soln and forms an osazone with phenylhydrazine. The exts of cholera and clear vibrios gave a strong pptn reaction with anti cholera serum, but are non antigenic unless a protein is added (serum or gelatin). In marked contrast, the ext of cloudy vibrios was antigenic as such. This is possibly due to its high N content (4.4%), which indicates impurities. II FAGLE

The immunological behavior of normal serum. III The appearance of complement. I FREIDBERGER AND J GURWITZ *Z. Immunitäts* 68, 351-63(1930).—The complement titer of newborn guinea pigs is usually about the same as that of the mother. The fetus of a pregnant guinea pig contains very little complement until immediately before birth, when the complement titer suddenly rises to approximate that of the mother. II FAGLE

Quantitative conditions in the formation of antibodies for lipoids by combination-immunization. FRANZ HEIMANN AND ELISABETH WEIL *Z. Immunitäts* 68, 403-8(1930).—Very small quantities of pig serum, as little as 0.005 cc, can be used in conjunction with lecithin as a "Schlepper antigen," to induce the formation of antibodies for the lecithin in rabbits. Similarly, 1-1000 serum used in conjunction with an EtOH rabbit heart ext suffices to induce the formation of antibodies for the rabbit lipoids. The more serum that is used, the less of the rabbit heart ext need be used to obtain antibodies. II FAGLE

The suitability of chemospecific antigen preparations for the formation of antibodies to lipoids. G L SELTER *Z. Immunitäts* 68, 409-27(1930).—The antisera obtained from rabbits following the injection of combinations of beef heart serum and diazotized stoxyl or melanilic acid give complement fixation with common lipoids, including lecithin, beef lipoids and cholesterol. A simple mixt of the diazonium compd and the serum possesses this antigenic activity. The antibody against the lipoids is contained entirely in the globulin fraction as obtained by acidification of the serum; the chemospecific antibody remains in the supernatant "albumen" soln. Unlike normal serum this diazonium serum mixt added to an EtOH ext of horse kidneys (Hersman antigen) does not make the latter antigenic. II FAGLE

The identity of animal-hair allergens (horse, cat and dog hair allergens). W. STORM VAN LEEUWEN *Z. Immunitäts* 68, 427-33(1930), cf. *C. A.* 24, 5844.—The allergens prepd from dog, horse and cat skin scales and hair are different. The skin sensitivity of human beings to these substances is not detd by a common factor, yet desensitization of a given skin area with one of these substances frequently brings about a concomitant desensitization to the other 2. This may possibly be due to the fact that the repeated reaction uses up some tissue component necessary for the skin reaction. II FAGLE

Comparative studies in the isoagglutinins in the blood of human beings and of the pig. S SCHIRMER, W KAYSER AND A KARNITZER *Z. Immunitäts* 68, 437-49(1930).—The isoagglutinin α (anti-A) of human sera corresponds to that of the pig. Pig serum contains cold agglutinins which make comparison difficult. II FAGLE

The heat susceptibility of complement fixation and flocculation reactions with syphilitic sera. MEINARD ZIEGLI *Z. Immunitäts* 68, 450-9(1930).—II syphilitic sera are heated at 60-68° for 5 hrs before testing, the Wassermann and Kahn reactions are equally affected in 21-90% and the Kohn reaction is more thermostable in 59% and more labile in 10.7%. Z. considers that this indicates a certain independence of the 2 types of test. II FAGLE

The heat susceptibility of complement fixation and flocculation reactions with syphilitic spinal fluids. MURPHY / *Clinical J. Immunology* 68, 479-65 (1930) — The Kahn reaction is more heat resistant than the Wassermann in spinal fluids. II. E.

The role of protective substances of the blood in the critical termination of attacks of fever and for the development of immunity in inoculated recurrent fever. I. FLAUT AND C. GIANOW / *Immunology* 68, 491-501 (1930) — The formation of serum antibodies in patients inoculated with *Spirillum* is not related to the bouts of fever, which are the clinical manifestation of the paroxysmal multiplication of spirochetes in the blood. They may, however, play a part in the final cessation of a fever and the death of the spirochetes. The immunity against re-infection for years following the inoculation is not deted by circulating antibodies, as shown by the absence of protective antibodies for mice. II. I. AGLR

Physical and chemical studies of the virus of hoof and mouth disease. I. Isoelectric point and adsorption. K. POPPE AND G. BUSCH / *Immunology* 68, 510-8 (1930); cf. C. A. 25, 2181 — Three strains of hoof and mouth disease virus had cataphoretic isoelectric points at pH 8.0, 7.6 and 7.8, resp. The virus is strongly adsorbed by 1 g of charcoal at pH 6.5-8.4, so strongly that neither the adsorbent or the supernatant fluid is infectious for guinea pigs. This adsorption is weakest near the isoelectric point of the virus. At this reaction, the virus may not be completely adsorbed; moreover, when the charcoal is injected into guinea pigs, there may be a disson of the adsorbed virus, with infection of the animal. Kaolin does not adsorb the virus. II. I. AGLR

The duration of strain-specific antibodies to trypanosomes in rabbits and the significance of the antibody content for the outcome of homologous reinfections. I. HARTMANN / *Immunology* 68, 518-30 (1930) — Using the Nagana-Prowazek strain of trypanosomes, II finds that protective antibodies for mice persist in the blood for 17-31 weeks after the therapeutic sterilization of the infected rabbit with neocarphenamine; they disappear in a few days from the blood of a passively immunized animal. Re-infection with a homologous strain, with the appearance of a chancre, is successful only if the circulating antibody is <1-2 units. Unlike syphilis, no infection without symptoms could be demonstrated, as shown by the injection of organ emulsions of suspected animals into normal rabbits. II. E. AGLR

Studies on blood volume. II. Influence of anaphylactic shock on blood volume. ISYAN S. WENT AND CECIL KENT DRINKER. *Magyar Orvos Arch.* 31, 54-8 (1930) — Expts on guinea pigs showed that blood vol. of animals did not change during anaphylactic shock. Cf. C. A. 23, 3942. S. S. DE L'VALY

Ratio of sensitizing capacity of anti-anthrax immune sera to their content of precipitin and complement-fixing antibodies. JÓZSEF TOMCSIK. *Magyar Orvos Arch.* 31, 211-9 (1930), cf. C. A. 24, 2001 — Purified ext. of encapsulated anthrax bacilli was prep'd according to Toennissen. Dilm. of the ext. to 1:1,000,000 gave a ppt with immune horse serum. The reaction on complement fixing bodies was neg. Immune rabbit sera gave both ppt. and powerful complement fixation in the presence of purified ext. Passive sensitization of guinea pigs could successfully be demonstrated with rabbit sera, being active in complement fixation. The passive sensitizing property of immune sera is approx. parallel to their content in complement fixing antibodies. S. S. DE L'VALY

Experiments to produce pure serum hemolysin. LAJOS (LUDWIG) GÖZONY AND FERENC (FRANZ) HOFFENKICH. *Monatsh. f. Chem. u. Physik* 46, 114-25 (in German 126) (1929), cf. C. A. 23, 5506 — Enzyme toxins which prevent the effect of lipase also decrease the effect of hemolysin. Methods of lipase production were used for the production of hemolysin since they show some kind of parallelism. Such hemolysin on purification was 40 times more active than the dried serum from which it was produced. Pure hemolysin is cryst. and develops acid from blood corpuscle lipoids and cannot be identified with pancreas lipase. S. S. DE L'VALY

Toxin-producing capacities of recently isolated strains of diphtheria bacilli. J. M. NEILL, R. C. AVERY, L. V. RICHARDSON, J. Y. SUGG AND B. E. KANE. *Am. J. Hyg.* 11, 399-403 (1930). *Physiol. Abstracts* 15, 555 — Strains of diphtheria bacilli of diverse origin had wide variations in toxin producing capacities. A large no. of the filtrates were also tested for their content of the heat stable substance responsible for the "pseudo" reaction of the Schick test, and also for the content of the heat stable constituent responsible for the immediate skin reaction of the diphtheria hypersensitive person. The filtrates varied in their content of these heat stable substances, and the amt. present bore no relation to the content of toxin. C. G.

The behavior of certain filterable viruses when subjected to cataphoresis. C. V. NATARAJAN AND R. R. HYDE. *Am. J. Hyg.* 11, 652-67 (1930). *Physiol. Abstracts* 15,

553—The myxoma virus of rabbits is electroneg. In the range of 4.0 to 9.3. Below pH 4.0 the virus is quickly inactivated. Its behavior in an elec. current as a negatively charged particle agrees with the fact that it is readily adsorbed by basic substances. The virus of epithelioma contagiosum behaves amphoterically, moving to both anode and cathode in an elec. field between a pH range of 6.4 to 9.3. The encephalitic virus (which had been passed through a series of rabbits by intracranial inoculation) proved to be negatively charged from pH 7.6 to 8.9. Below pH 7.6 it is pos. Different strains of bacteriophage were tested, and varying results obtained. All attempts to use the method of cataphoresis as a means of sepg. the viruses from assoc. proteins failed.

G G

The chlorine-sodium ratio of blood serum in edema from Bright's disease. M. LAUDAT AND A. GRANDSIRE. *Compt rend soc biol* 103, 893-5(1930), *Physiol Abstracts* 15, 305.—In edema from Bright's disease there is a marked increase of the Cl/Na ratio of the serum. This is maintained in some cases as the edema disappears, but a relation exists between the hydration of the tissues and the increase of the ratio.

G G

The chlorine-sodium ratio of blood serum in Bright's disease. M. LAUDAT AND A. GRANDSIRE. *Compt rend soc biol* 103, 1212-6(1930), *Physiol Abstracts* 15, 378.—The Cl/Na ratio varies in cases of N retention in Bright's disease. The increased concn of blood urea does not seem to intervene directly in the variations, which depend on assoc. factors. Prolonged retention of water raises the ratio, vomiting, on the contrary, diminishes it.

G G

Natural hemolysins of human serum. J. JADIN. *Compt rend soc biol* 103, 951-3(1930), *Physiol Abstracts* 15, 242.—The content of hemolysins in the blood of the different human groups was detd. for sheep corpuscles and for those of guinea pigs. The greater no. of sera contain both hemolysins, but those of group B are the most active.

G G

The serum-antitoxin complex does not carry the alexin through the placenta. L. NATTAN-LARRIER AND L. RICHARD. *Compt rend soc biol* 103, 1084-6(1930), *Physiol Abstracts* 15, 413, cf C A 25, 1574.—Pregnant guinea pigs were immunized passively and actively against diphtheria and tetanus toxins. The serum antitoxin complex filters through the placenta, but it does not contain the alexin. In one instance the alexin was found in the fetal blood, but this fetus was at term, and so it is believed that the alexin was of fetal origin.

G G

Antigenic properties of an extract demonstrated by the Prusnitz-Kustner method. P. VALLERY-RADOT. *Compt rend soc biol* 103, 1207-8(1930), *Physiol Abstracts* 15, 413.—When it is desired to ascertain if an ext. is actively antigenic, it is enough to inject it into the skin of a human subject at the site of an injection, made 21 hrs. before, of serum from an individual whose specific sensibility to the antigen is well established. The proof of activity of the ext. is given by a marked local reaction. This test provides the best method of demonstrating that an antigenic ext. has the required properties for exptl. research on anaphylaxis.

G G

Action of globulins on the development of transplantable lymphosarcoma of the mouse. J. RICARDO MEYER. *Compt rend soc biol* 103, 1322-3(1930), *Physiol Abstracts* 15, 415.—Globulin pptd. from horse serum with $(NH_4)_2SO_4$ was injected into mice, which were subsequently grafted with a lymphosarcoma, while controls were grafted without injection of globulin. The percentage of pos. results in the former was 90, in the latter only 50, while the dimensions of the growth were only about $1/2$ in the non treated mice. Grafting expts. performed after the injection of U nitrate (which causes an increase of globulin in serum) show that the action of the globulin in favoring cancer growth is not a direct nutritive one, but that it is secondary to interference with the reticulo endothelial system.

G G

Distribution of antibodies in the rabbit organism. O. G. BIER. *Compt rend soc. biol* 103, 1324-30(1930), *Physiol Abstracts* 15, 413.—In rabbits deprived of their kidneys, sensitizing substances injected were maintained at the same concn. in the blood during 44 hrs.

G G

Diffusibility of potassium in normal human muscle and in the condition of pyramidal and extra-pyramidal contracture. O. SAGER AND E. ROTTE. *Compt rend soc. biol* 103, 1373-4(1930), *Physiol Abstracts* 15, 340.—No K is fixed to the colloids in human muscle. All the K diffuses out in time into the surrounding medium. The speed of diffusion of K is not the same for normal and pathol. muscle; it differs even in the different pathol. conditions which are detd. by differences of the colloidal state of the fiber-sarcoplasm phases.

G G

Blood pH in cases of cutaneous epithelioma. R. REDING AND A. SLOSSE. *Compt.*

rend soc biol 104, 124-7(1930), *Physiol Abstracts* 15, 415—Variations of the pH of blood in different types of tumor have been reported. In the present investigation 23 cases of skin cancer were examd. The results are contradictory to those of Jaumain. The pH in the 23 cases varied from 7.41 to 7.47. There is no alteration in pH in cancer of slow evolution. In case of marked activity it is natural to expect alteration of the physicochem. properties of the blood. G. G.

Excess of blood potassium in cancer. A. RAMOND AND E. CANTEGRIL. *Compt rend soc biol* 104, 293-4(1930), *Physiol Abstracts* 15, 457, cf. *C. A.* 25, 1575—Excess of K has been found in the blood of 83% of cases of cancer investigated (48). The cases in which it is least marked are those with facial cancer. These observations correspond with those made on the K content of cancerous growths. The content is greatest when proliferative activity is most marked. G. G.

Urinary excretion in normal and nephritic dogs on variable diet. P. GENAUD. *Compt rend soc biol* 104, 545-9(1930), *Physiol Abstracts* 15, 457—Glucosuria always exists in U intoxication and is unaffected by change of food. In the nephritic dog chloride is not retained. The output of urea is greater in the normal animal than in the nephritic animal on a similar diet. NH_3 is const. in amt. and is independent of food in nephritis. The excretion of creatinine is not influenced by food, that of creatine is greater in nephritis. Uric acid is about the same in amt. in both cases. G. G.

Non-sugar reducing substances of human blood in pathological conditions. FAEDA K. HERRERT AND MARGHERITA C. BOURNE. *Biochem J* 24, 1787-93(1930), cf. *C. A.* 24, 5354—The concn. of the non-diffusible reducing substances is shown to vary with the hematocrit reading for the corpuscle vol. BENJAMIN HARROW.

Alimentary intoxication. F. ELMER JOHNSON. *Arch Pediatrics* 48, 194-201 (1931)—In a case of alimentary intoxication of infancy, the blood findings were normal. In another case, the blood analysis suggested both acidosis and uremia (80 mg. total non protein N, 161 mg. dextrose, 7 mg. P, 500 mg. chloride per 100 cc. of blood; and CO_2 -combining power 12.6% by vol.) JOSEPH S. HEPBURN.

Flocculation experiments with variola and vaccinia virus. LEON C. HAVENS AND CATHERINE R. MAYFIELD. *Am J Pub Health* 21, 329-43(1931)—Fppl evidence indicates that the intradermal or intravenous inoculation of vaccinia virus in rabbits is followed by the appearance of sp. flocculating properties in their sera. Vaccination of the human subject stimulates the production of precipitins against vaccinia virus. The serum of 4 persons acquired this power to flocculate the virus as early as the 10th day following vaccination. J. A. KENNEDY.

The formation of phospholipides during the autolysis of normal and neoplastic tissue. A. H. ROFFO AND L. M. CORREA. *Bull soc. chim. biol* 12, 1247-54(1930)—Normal dog and rat livers and spleens were autolyzed in NaF soln. and compared with the same tissues from animals with cancer. Differences were found in the quantities and rates of hydrolysis of P and fatty acid compds. C. G. KING.

Observations concerning the causative agent of a chicken tumor. JAMES B. MURPHY, O. M. HILMER, ALBERT CLAUDE AND ERNEST STUM. *Science* 73, 266-8 (1931)—A report of addnl. observation on the properties of a filtrable agent causing chicken tumor. The agent of Chicken Tumor 1, a spindle-cell sarcoma, is selectively adsorbed and fixed by certain mesodermal tissues from susceptible animals. The plotted curve of the amt. of ultra-violet light of selected wave lengths required to inactivate the tumor agent shows a significant qual. and quant. variation from the curves for bacteria, typical viruses and bacteriophage. The tumor producing activity of the filtrates can be pptd. out with a protein fraction and purified. The steps in the purification of the agent are outlined and the evidence of an inhibiting principle in the chicken tumor is discussed. B. S. LEVINE.

Individual differences in human blood. KARL LANDSTEINER. *Science* 73, 403-9 (1931)—Precipitation and agglutination are reviewed with particular reference to blood grouping as a possible indicator of individual differences in human blood. In man there are numerous individual blood differences already demonstrated, and undoubtedly there exist still others which have not yet been established. Whether actually each individual blood possesses a special quality or how frequently there is complete correspondence with the blood of others, cannot be definitely stated at present. B. S. L.

A soluble specific carbohydrate of ragweed pollen. J. H. BLACK. *J. Allergy* 2, 161-3(1931)—Evidence is presented showing the presence of a complex carbohydrate in ragweed pollen. This substance reacts specifically in the skin and nasal mucosa of ragweed sensitive individuals. JULIAN H. LEWIS.

The stabilization temperature of sera in the Meimecke and Wassermann reactions. JOSEPH HOHN. *Z. Immunitäts* 67, 30-8(1930)—Heating sera at 51° for 4 min. does not

affect the Meinecke clearing reaction and makes the Wassermann reaction sharper. If heated in amts less than 0.5 cc. or over 52° unspecific reactions with the Meinecke method are produced. A quick Meinecke test in which the centrifuge is used is described. JULIAN H. LEWIS

A common antigen in human cells and Shiga bacilli. M. EISLER. *Z. Immunitäts* 67, 38-43(1930).—In Shiga dysentery bacilli and in cells of human organs, including carcinoma cells but not serum, is a common antigen which produces in goats, but not in other animals, an antiserum that agglutinates the Shiga bacilli and human red cells. This antiserum does not hemolyze human red cells and does not contain Forssman antibodies. JULIAN H. LEWIS

The value of the citochol reaction in the diagnosis of syphilis. D. P. BOROVSKAYA AND S. D. ORLOVA. *Z. Immunitäts* 67, 63-6(1930).—The citochol reaction of Sachs and Witebsky is found to be a sensitive, sp., simple and const. test for syphilis. J. H. L.

The normal antibodies of rabbits at different ages. E. FRIEDBERGER AND D. GAJZÁGÓ. *Z. Immunitäts* 67, 67-74(1930).—Rabbits first form normal sheep hemolysins about the 75th day after birth, this is comparable to man in whom these antibodies are formed during the first year. Newborn rabbits have sheep hemolysins in their sera which may have a higher titer than either of the parents. These antibodies disappear completely in 8-14 days. Blood from the umbilical cord of human babies contains no hemolysins, but because of technical difficulties it is not known if the antibodies are in the general circulation during the first day of life. JULIAN H. LEWIS

Protein anaphylaxis in tuberculous guinea pigs as compared to non-tuberculous controls. E. FRIEDBERGER AND D. GAJZÁGÓ. *Z. Immunitäts* 67, 75-8(1930).—Tuberculous guinea pigs sensitized with sheep serum withstand larger doses of the serum on re-injection than do controls. The difference is not marked but is demonstrable. JULIAN H. LEWIS

The immune bodies in fractions of antiserum for foot and mouth disease. The increase in titer of immune serum after parenteral injection of serum fractions. P. V. GARA AND K. TRAUTWEIN. *Z. Immunitäts* 67, 102-11(1930).—The protective action of antiserum for foot and mouth disease resides in the typhobic euglobulin fraction. This is true of antiserum for the different types of virus and from calves and guinea pigs. It was not possible to increase the antibody content of convalescent guinea pigs by injecting the guinea pigs with immune serum fractions before infection. J. H. L.

The production of lipid antisera by injections of organ suspensions. FELIPE MORAN. *Z. Immunitäts* 67, 115-25(1930).—Injections into rabbits of a suspension of liver produce an antiserum specific for liver lipids but this specificity is limited to within the species. Injections of suspensions of thyroid produce a similar but less-marked sp. serum. Other organs in suspension produce ubiquitous lipid antibodies. With the exception of the lens, suspensions and alc. exts. of organs of the eye do not produce organ-sp antibodies. Brain-sp antisera react with alc. exts. of the retina. JULIAN H. LEWIS

The antagonism between group specific antigen and group specific antibody in rabbit blood. K. HARA. *Z. Immunitäts* 67, 125-36(1931).—Complement-fixation tests using anti group A human red cell sera as antiserum and native rabbit serum or its alc. ext. as antigen show that the serum of rabbits contains group A antigen if the red cells of these rabbits lack the agglutinin of Group A. JULIAN H. LEWIS

The group differentiation of fetal membranes. R. S. CHERIKOVER AND O. M. SEMZOVA. *Z. Immunitäts* 67, 240-50(1930).—Neither fetal tissues nor the uterine decidua contain group-sp antigens. The antigen found by Oettingen and Witebsky (*Munch. med. Wochschr.* No 3, (1928)) in the decidua was the heterogenetic Forssman antigen. JULIAN H. LEWIS

The action of carbon dioxide on the coagulation and the complement action of plasma. HANS J. FUCHS. *Z. Immunitäts* 67, 266-71(1930).—Increase of CO₂ of plasma increases its complement action and coagulation speed as long as the reaction remains alk. or neutral. These properties are decreased when the reaction becomes acid. JULIAN H. LEWIS

Complement and anticomplement. HANS J. FUCHS. *Z. Immunitäts* 67, 272-85(1930), cf. C. A. 24, 422.—Complement is considered as a series of chain reactions, one component of which is the middle piece that is identical with prothrombin. The middle piece does not occur in the flowing blood, nor is it completely free in plasma and serum but is more or less firmly bound to its antagonist or stabilizer, antiprothrombin. There are 3 classes of anticomplementary action. The first includes the antiprothrombin excess of hemophilic blood, peptone blood of dogs and heparinized blood; serum inactivated at 56°; high salt concns. The second class includes colloids that adsorb the

antithrombin prothrombin complex, such as $\text{Ca}_3(\text{PO}_4)_2$, $\text{Mg}(\text{OH})_2$, BaSO_4 , CaC_2O_4 , org sp ppts and bacteria. The third class are the phosphatides, especially cephalin, which neutralizes antiprothrombin, permitting the prothrombin to be converted into thrombin which has no complement action. JULIAN H LEWIS

Complement-fixation antibodies against alcoholic extracts of carcinoma in patients with carcinoma and in pregnancy. L HIASERFELD, W HALPFA, M FLOKSTRUPFF, AND J KOŁODZIEJSKI. *Z Immunität* 67, 286-318(1930).—By using an alc ext of carcinomas fortified with cholesterol carcinoma can be diagnosed in 50-60% of cases with the complement fixation reaction. In 5% of the controls with neg Wassermanns the reaction is unspecific. Controls with pos Wassermanns give also pos reactions with the alc ext of carcinomas. Some of the carcinoma patients give pos Wassermann reactions. In pregnancy 60% of the sera reacts with the alc ext of carcinomas. JULIAN H LEWIS

Demonstration of impedin in bacteria. M YOSHITOMI. *Z Immunität* 67, 353-8(1930).—Typhoid bacilli contain a substance called impedin that inhibits the production of agglutinins when they are used as antigen and that is responsible for the toxicity of vaccines. While impedin is present only in the bodies of washed bacteria, they are gradually dissolved in the medium of a suspension on standing. Impedin is destroyed by boiling and the resulting prepn is called cocto-vaccine. J H L

The production of cocto-vaccine. R TORIKATA AND M. YOSHITOMI. *Z Immunität* 67, 358-67(1930).—The optimum period of boiling for the production of cocto-vaccine in which impedin, a toxic and antagonistic substance, is destroyed is 30-60 min. Cocto-vaccine is also better adapted for use as an *in vitro* antigen. JULIAN H LEWIS

Complement and media of salt solutions. JOHANN SCHURFAT, *Z Immunität* 67, 367-9(1930).—Isotonic zones of various salts are given. It is shown that other salts than NaCl may serve as media for complement. Bromides and iodides enhance the action of complement. JULIAN H. LEWIS

The rate of disappearance of proteins and lipoids from the peritoneal cavity of guinea pigs. JOHANN SCHURFAT. *Z Immunität* 67, 370-2(1930).—By means of chem and serological tests it is found that egg white, cholesterol and lecithin injected into the peritoneal cavity of guinea pigs disappear completely in 24 hrs. J H. L.

Conditions and methods for demonstrating carcinoma-specific antigens. HERMANN LEHMANN-PACIUS AND TADAO TODA. *Z Immunität* 67, 373-85(1930).—Carcinoma suspensions heated 0.5 hr at 100° produce in rabbits a serum highly sp for carcinoma cells. This specificity depends on lipid antibodies. Individual sp factors or the blood group of the antigen used for immunization have no effect on the specificity of the serum. Antisera against boiled suspensions of carcinomas of group A contain, in addn to carcinoma sp antibodies, also antibodies sp for group A. J H. L.

Experimental studies on the properties of diphtheria anatoxin. M P. ISABOLINSKI AND V. I. GITOVICH. *Z Immunität* 67, 441-7(1930).—Guinea pigs injected with anatoxin were immune to toxin as well as to pure cultures of diphtheria bacilli. No difference in the antigenic action of the usual anatoxin and Ramon's total anatoxin could be found. JULIAN H LEWIS

Experimental studies on the Abderhalden reaction. M LUJAK, N KORSANOW AND A ROSENBLATT. *Z Immunität* 67, 462-74(1930).—Organs of dogs or sheep transplanted into rabbits did not produce positive Abderhalden reactions of the rabbit sera with the organs. Pregnancy in rabbits and guinea pigs could not be diagnosed with the Abderhalden reaction. The sera of these animals are not suitable for the test because most of the sera without antigen give positive reactions. Positive reactions could not be obtained in dogs by immunizing with egg white or foreign organ emulsions but were obtained by transplanting foreign organs or after auto-immunization by injuring certain organs of the dog. These positive reactions, however, were not specific. J H L.

Antibodies for leucocytes. E WITENSKY AND K KOMIYA. *Z Immunität* 67, 480-96(1930).—Rabbits immunized with guinea pig leucocytes produce Forssman antibodies and antibodies that react with alc. exts of guinea pig leucocytes and spleen. When immunized with rat leucocytes species sp lipid antibodies are formed without any organ differentiation. The specificity of organ sp antisera is increased by adding lecithin to the organ exts while the nonsp quota is inhibited. The organ sp lipid antibodies of leucocyte antiserum are usually limited to the species. Guinea pig leucocyte antisera react exclusively with the homologous exts of leucocytes and spleen but not with those of blood. On the other hand, antisera for guinea pig red cells react also with exts of guinea pig leucocytes and spleen. Boiled suspensions of leucocytes react with the sp antisera just as well as with native suspensions. Leucocyte antisera do not

cause agglutination or lysis of leucocytes. Rabbit leucocytes react with leucocyte antiserum obtained from the rabbit. JULIAN H. LEWIS

Preservation of complement with sodium acetate and boric acid. CURT SONNENSCHEIN. *Z. Immunitäts* 67, 512-6(1930).—Complement is better preserved with 10% NaOAc and 4% H_3BO_3 added in substance than with 10% NaCl and 4% H_3BO_3 (C. A. 24, 5853). JULIAN H. LEWIS

Influence of temperature on the sensitivity and specificity of the Wassermann reaction with different types of antigen. GÁBOR OLÁSI. *Z. Immunitäts* 67, 517-38 (1930).—The sensitivity of the Wassermann reaction is increased if the serum complement antigen mixt is incubated at low temps. There is, however, a considerable reduction in specificity of the reaction. Many sera react better with exts contg no cholesterol when the test is made at low temps. JULIAN H. LEWIS

The action of carbon dioxide on complement. HANS J. FUCHS. *Z. Immunitäts* 69, 51-7(1930). cf C. A. 24, 5843.—Blood corpuscles with their O displaced with CO_2 adsorb the middle piece of complement completely from plasma. Hemolysis does not result. The plasma treated with the red corpuscles shows a loss of spontaneous coagulability as well as of complement activity. This indicates the identity of complement middle piece and prothrombin. JULIAN H. LEWIS

The serological relation of the flocculi and the fluid in flocculation tests for syphilis. D. P. BOROVSKAYA. *Z. Immunitäts* 69, 76-85(1930).—The flocculi produced in the different flocculation tests for syphilis contain antigen and antibody in various proportions. The supernatant fluid in the Meinel reaction reacts like luetic serum while that in the Kahn and citochol reactions reacts like antigen. JULIAN H. LEWIS

Complement fixation with snake venom immune serum. VITAL B. FILHO. *Z. Immunitäts* 69, 126-33(1930).—The snake venoms differ greatly in their ability to inhibit sp. hemolysis. In prep. the proper concn for use as antigen in complement fixation expts. this had to be taken in account. Complement fixation is best obtained when fresh unheated immune serum is used. Normal horse serum showed a peculiar reaction in that it produces complete fixation with small doses of certain venoms but none with larger doses. There is a strong complement fixation between *Crotalus terrificus* toxin and antitoxin but none between the toxins of the *Lachesis* and their antitoxin. There was an absence of reaction between the toxin and antitoxins of the *Crotalus* group and those of the *Lachesis* group. No relation could be established between the complement-fixing power and the neutralizing strength of the *Crotalus* antitoxin. JULIAN H. LEWIS

The chemical nature of the heterogenetic antigen in Shiga bacilli. KURT MEYER. *Z. Immunitäts* 68, 98-108(1930).—By treating a suspension of Shiga bacilli with anti-formin and pptg. with alc. in alk. reaction, a carbohydrate is obtained comprising approx. 0.5 of the total heterogenetic antibody in the original bacilli. H. EAGLE

The chemical nature of the heterogenetic antigen of the Shiga bacillus. KURT MEYER. *Z. Immunitäts* 69, 134-45(1930), cf preceding abstr.—The heterogenetic antigen of the Shiga bacillus is a carbohydrate and is different antigenically but not chemically from the carbohydrate that carries species specificity. *Ibid.* 499-517(1931).—The species specificity and heterogenetic activities of Shiga bacilli reside in the carbohydrate isolated from these organisms. Sepn. by sp. pptn. and adsorption of the carbohydrate into fractions corresponding to each of these activities could not be accomplished. This indicates that the 2 properties are firmly bound and that they represent 2 reactive groups of a single substance. JULIAN H. LEWIS

Respiration of bacteria and immune reactions. Experiments with *B. pyocyaneus*. GYULA SURANYI AND JOSEF V. PALOCZY. *Z. Immunitäts* 69, 161-8(1930).—Agglutination with immune serum lowers the consumption of O by *B. pyocyaneus*. This effect is not the effect of agglutination itself but is due to a decrease of respiratory surface. *Immune serum and complement producing bacteriolysis increase O consumption at first, after which there is a complete cessation of respiration.* JULIAN H. LEWIS

An unusually high degree of anaphylactic sensitization. Acute lethal anaphylactic shock after subcutaneous injection of small doses of antigen. R. DÖRR AND S. SEIDENBERG. *Z. Immunitäts* 69, 169-79(1931).—In hereditary anaphylactic sensitization there is a much higher degree of sensitization in guinea pigs than after the usual active sensitization. Animals react fatally to subcutaneous injections of 0.2-2.0 cc. doses of serum antigen with a latent period of 15-20 min. when after other forms of sensitization, except under unusual conditions, there is little or no reaction after subcutaneous injections. JULIAN H. LEWIS

The use of alcoholic extracts of meningococci for complement fixation and titration of meningococcus antiserum. H. SACHS. *Z. Immunitäts* 69, 221-39(1930).—The alc. exts. of meningococci give complement-fixation reactions with antimeningococcus serum.

and can be used to titrate the antiserum. *Meningococcus* antiserum contains, therefore, sp. lipid antibodies but it also contains non-sp. lipid antibodies since it gives positive Wassermann and citochol reactions. The activity of alc. exts. of meningococci is increased by the addition of lecithin. JULIAN H. LEWIS

Complement fixation and flocculation with diphtheria antiserum. F. HÖRING. *Z. Immunitäts* 69, 244-60(1930).—Aic. exts. of diphtheria bacilli and diphtheria toxin give strong complement fixation reactions with diphtheria antiserum. These reactions are intensified by the addn. of a small quantity of lecithin to the antigens. The flocculation of toxin and agglutination of bacilli by antitoxin are facilitated by the addn. of citochol. Treatment of antiserum with diphtheria bacilli removes the complement-fixing antibodies for the alc. ext. of diphtheria bacilli but the flocculating antibodies remain. The enhancing effect of lecithin can be substituted with cholesterolized and plain alc. exts. of organs. With none of these reactions is there a complete parallelism with the antitoxin content of the antiserum. JULIAN H. LEWIS

The role of cholesterol in the activation of lipid antigens. G. F. DE GAETANI. *Z. Immunitäts* 69, 277-97(1930).—Forssman antigens that are poor in lipoids are activated by the addn. of cholesterol. This effect is due not to cholesterol antibodies but to some physicochem. influence on the antigenic substrate. JULIAN H. LEWIS

The chemical explanation of immune hemolysis. L. JARNO AND L. SKRANYL. *Z. Immunitäts* 69, 298-304(1930), cf. *C. A.* 24, 420.—A study of the enhancing effect of Na cholate and various amino acids on hemolysis and an analysis of inhibition of hemolysis by KCN and caffeine. Na benzoate lead to the conclusion that normal and immune hemolysis are to be considered as compds. of chole. acid coag. complexes of amino acids. JULIAN H. LEWIS

Is the diphtheria toxin-antitoxin mixture more toxic for tuberculous than for normal guinea pigs? E. FRINGSBERGER AND O. ANDERSON. *Z. Immunitäts* 69, 313-7(1931).—Tuberculous guinea pigs show no increased sensitivity to underneutralized mixts. of diphtheria toxin-antitoxin mixts. On the other hand they show an increased resistance to anaphylaxis and to other infections. JULIAN H. LEWIS

The consumption of complement in hemolysis. HANS J. FUCHS. *Z. Immunitäts* 69, 330-42(1931), cf. *C. A.* 24, 418.—In hemolysis of sensitized red cells by complement there occur, side by side, 2 different reactions. One is the use of prothrombin as the mid piece in complement activity and the other is the conversion of prothrombin into thrombin. The speed of these 2 reactions depends on the conditions existing in the mixt. If the red cells are weakly sensitized the part of the prothrombin not used for complement is changed into thrombin through the action of the cytochrome-phosphatide liberated by hemolysis and of Ca in the serum. The prothrombin undergoing this change can no longer be utilized for complement. This explains the fact that blood fluids contg. Ca lose their complement activity sooner than those that are free from Ca. JULIAN H. LEWIS

Titration and detection of snake venom antisera by intracutaneous tests into guinea pigs. R. OTTO. *Z. Immunitäts* 69, 368-78(1931).—By injecting mixts. of snake venom and antiserum intracutaneously into guinea pigs the neutralizing power of the antiserum can be titrated. Of 117 patients only 9 were found to have antivenom in their serum. JULIAN H. LEWIS

Comparison of the antigenic action and toxicity of scarlet-fever toxins modified by heat and by sodium ricinolate. CH. CH. WANG. *Z. Immunitäts* 69, 395-401(1931). The immunizing action of scarlet fever cocto-antigen. *Ibid.* 402-7. Vaccination with scarlet-fever cocto-antigen. K. SAIRO. *Ibid.* 408-12. Both heat and 1% Na ricinolate reduce the toxicity of the Dick scarlet fever toxin, but Na ricinolate lowers the antigenic action, while heat raises it as proved experimentally and with patients. J. H. L.

The new flocculation tests, especially the Kahn test. CARL SCHLESSELMANN. *Z. Immunitäts* 69, 484-98(1931).—The Meenecke clearing reaction is slightly sharper and more sp. than the Kahn test and is considered the best flocculation test so far devised. JULIAN H. LEWIS

The thermic effect of death and hemolysis. V. V. LEFESHKIN. *Am. J. Physiol.* 95, 473-80(1930).—The hemolysis or the coagulation of protoplasm of red corpuscles is accompanied by the production of heat to the amt. of 2.1-2.4 cal. per g. of dry substance of corpuscles. Death is accompanied by the decompn. of some very unstable compd., probably in this case compds. of hemoglobin. J. F. LYMAN

Experimental arteriosclerosis in the rat. M. SKEENEY AND ERMA SMITH. *Am. J. Physiol.* 95, 620-5(1930).—Arteriosclerosis could not be produced in adult rats (1) by feeding for 7 months a diet contg. 5% NaCl and 10% NaHCO₃ or (2) by const. intensive infection for 7 months by weekly injections of *Staphylococcus aureus*. The ingestion of

50 cc. of viosterol over a period of 25 days produced definite arteriosclerosis with densely calcified aortae. In one case the renal blood vessels appeared to be calcified.

J. F. LYMAN
Acidity of the gastric contents of normal, cretin and hyperthyroid rabbits. M. F. GREEN AND M. M. KUNDE. *Am J Physiol* 95, 628-9(1930).—The total HCl of the gastric contents of normal and cretin rabbits is the same ranging from 0.2 to 0.56%. With exptl. hyperthyroidism, HCl of the gastric contents is depressed but not absent.

J. F. LYMAN
Diabetes insipidus. IV. HELEN BOURQUIN. *Am J Physiol* 96, 66-77(1931), cf *C A* 23, 3970.—Diabetes insipidus is not a deficiency phenomenon. The mamillary bodies, or centers in their vicinity, are essential to it and the polyuria is caused by a diuretic substance, or irritation of hypothalamic centers controlling salt and water metabolism.

J. F. LYMAN
Calcium shifts in experimental rickets. J. MORELLE. *Proc. Physiol. Soc., J Physiol* 70, xiii-xiv(1930).—In healing of rickets the Ca is at least partly provided (1) by the immediate neighborhood and (2) by the spongy trabeculae which constitute a reservoir of readily available Ca.

J. F. LYMAN
Detection of histamine-like substances in asthmatic sputa and experiments on their possible bacteriological origin. F. A. KNOTT AND G. H. ORIEL. *Proc. Physiol. Soc., J Physiol* 70, xxxi(1930).—Bioassay indicated the presence of histamine in several sputa from asthma patients whose bronchial plugs and mucopurulent masses contained Gram-negative bacilli, apparently allied to the group of Friedländer's bacillus. Cultures of these organisms also gave histamine-like effects. In bronchial infection with these bacilli, therefore, one may have one cause of sp. locally produced asthma.

J. F. LYMAN
The occurrence of a non-heat-coagulable protein in normal and pathological urines. G. H. ORIEL. *Proc. Physiol. Soc., J Physiol* 70, xxxvi(1930).—A non-heat-coagulable protein can be sepd from the urines of persons suffering from bacterial invasions and especially in such cases of allergy as eczema, urticaria and asthma. The sepn can be made by adsorption on kaolin from acidified urine and elution by adding NaOH to the sepd. Kaolin to pH 8.0. It is pptd by catn with $(NH_4)_2SO_4$ and passes through a semi-permeable membrane with difficulty. Rabbits can be sensitized with it so that they show the Arthus phenomenon. By this technique it can be shown that the substance is not identical when obtained from persons suffering from different diseases. Persons suffering from allergic diseases are sensitive to this substance from their own urines and it is possible to desensitize them by minute doses at suitable intervals.

J. F. LYMAN
Starch digestion in man. A. F. HURST AND F. A. KNOTT. *Proc. Physiol. Soc., J Physiol* 70, xxxvi(1930).—A type of intestinal carbohydrate dyspepsia, characterized by the presence of undigested starch in the colon, is thought to be due to a failure of normal secretion of the digestive juice produced by the small intestine and a deficiency of intestinal diastase.

J. F. LYMAN
Some questions and problems of diabetes research including some evidence of a metabolic-physiological theory of diabetes mellitus. H. CHR. GEELMUYDEN. *Ergebnisse Physiol.* 31, 1-95(1931), cf *C A* 24, 5369.—"Insulin has no direct influence upon the oxidation of carbohydrates in the animal body. It fails to increase the total heat production and metabolism in the animal body. It can accelerate carbohydrate oxidation by providing intermediary compds. whose existence conditions carbohydrate oxidation such as muscle work (hypoglycemic convulsions), fat and glycogen synthesis. Furthermore insulin can promote counter reactions of the nervous-endocrine metabolic regulation by which the total oxidation and thereby that of carbohydrates can be accelerated positively or decreased. Such considerations probably explain the diverse findings of various authors in their studies of the effect of O_2 intake upon insulin. The basic reaction of insulin appears to produce a decrease in the formation of carbohydrates from non-carbohydrate materials, this in turn forces the body to burn preformed carbohydrates. This results in an increased R. Q. and a decreased O_2 intake. This basic type of insulin reaction is in accord with the concept of diabetes mellitus as a pathol. acceleration of sugar formation from non-carbohydrate materials."

C. M. MCCAY
Experimental investigation into the appearance of protein in bile. II. Hetero-albuminemia. T. MARUDA. *Japan J Gastroenterology* 3, 14-17(1931); cf *C A* 25, 328.—Egg albumin injected intravenously into rabbits is excreted into the bile by the parenchymatous cells of the liver. This was found after blocking the parenchymatous and stellate cells in turn with CCl_4 and Indian ink, resp. Rabbits whose livers were injured by CCl_4 displayed a decreased ability to excrete foreign protein in the bile.

C. M. MCCAY

Biochemical investigation of the blood in cases of experimental disturbance of liver function. II Liver function and protein metabolism. TSUNEJIRO SAWADA *Japan J Gastroenterology* 3, 34-47(1931), cf *C A* 25, 1282.—The livers of rabbits were injured by infection with *Schistosomum japonicum*, by ligation of the ductus choledochus or by P poisoning. Such injuries increase the amino acids and serum globulin of the blood and decrease the fibrinogen and serum albumin. C. M. McCAY

Jaundice in malarial diseases. Clinical and experimental investigations. I. Clinical research. HANFUSHI SAKET *Japan J Gastroenterology* 3, 46-53(1931).—The concn of bilirubin in the blood of patients inoculated with malaria was detd at regular time intervals during the course of the attacks. Malarial patients suffer from hyperbilirubinemia. The excretory function of the liver is impaired. This contributes to the jaundice as well as the abnormal increase in hemolysis. C. M. McCAY

The significance of the liver glycogen in the liver function. III. 1. The glucose test in cases where the liver glycogen is diminished. TANESABURO KUSUMAKI *Japan J Gastroenterology* 3, 51-9(1931), cf *C A* 25, 1277, 1278.—The blood sugar of rabbits was detd by the Hagedorn Jensen method at regular intervals after the injection of 20 cc. of 25% glucose intravenously. After fasting, or after the injection of adrenaline, strychnine or phorbium the blood sugar rises to higher levels and persists longer after glucose injections. Ligation of the bile duct has a similar effect. I believes such phenomena are the result of impaired liver function when its glycogen content is decreased. C. M. McCAY

Clinical and experimental studies on urobilin bodies. III. The urine-, bile- and blood-urobilin-bodies in cases of experimental hepatic disturbance. MASASAO OSHIMA *Japan J Gastroenterology* 3, 67-70(1931), cf *C A* 24, 4833.—A normal rabbit excretes 0.93-0.07 mg. of urobilin bodies per day in the urine and 80 times as much in the bile. Injury to the liver by either CHCl_3 or CCl_4 increases the excretion of the urobilin compds. In normal rabbit bile the proportion of urobilinogen to urobilin is 5:1. After hepatic damage the urobilinogen increases in the bile. C. M. McCAY

Autochthonous poisons of a lysoeythine nature in the brain. G. PICHINI *Boll soc ital biol sper* 5, 1131-8(1930)—Alec. exts. of human brain, normal and pathological, were treated with 1:10 to ppt. the so-called lysoeythins. The amt. of ppt. from the pathological brains was 5 to 10 times that from the normal brains. The senile dementia and the postencephalitic brains gave the most ppt. (0.4-0.5 cg. from 800 g. tissue). The hemolytic power of these ppts. was absent. When these substances were injected subdurally into rabbits there was an intensive neurotoxic reaction, whereas the material from the normal gave little or no reaction. The substances when injected subcutaneously into the ear of the rabbit produced hemorrhagic edema. These properties lead to the belief that the substances isolated from the pathological brains belong to the lysoeythins and are by products of the scission of one or more phosphatides contained in the nerve tissue. PETER MASUCCI

The nature of the period of incubation. I. NICO SERENI *Boll soc ital biol sper* 5, 1150-5(1930)—In most biologic phenomena the interval between the application of an initial stimulus and the response to the same stimulus is very brief, in anaphylaxis, the time between the sensitizing dose and the shock dose is much longer, regardless of the delicacy of the criteria chosen as the index. The long period of incubation, however, is a concept essentially bound to the symptomatic aspect of anaphylaxis and not to that strictly immunological. The period of incubation (clinical) of anaphylaxis is essentially the interval of time necessary for the sensitization of the organs of shock, by the antibodies produced in the organs of production, it consists of the sum of the times necessary for the antibodies to pass in sufficient amts. from the producing organs into the blood and from the blood into the various organs in which they are fixed. The antigen reaches the organs producing the antibodies more or less rapidly according to the route of introduction, and initiate immediately the production of antibodies. At this moment the true period of incubation comes to an end. In passive anaphylaxis, the antibodies have been formed by the producing organs of another animal but the role played by the organs of shock is entirely the same in that they do not play a part in the true immunological process. PETER MASUCCI

The inorganic sulfur of the blood in nephropathic cases. UGO DE MICHELIS. *Minerva med* 1931, 1, 318-24.—The serum was deproteinized by means of trichloroacetic acid. The filtrate was treated with BaCl_2 and the resulting ppt. was matched against a known standard BaSO_4 suspension. The inorg. S was detd. on 12 normal individuals and on 37 cases suffering from various renal lesions. The inorg. S in the latter group was markedly higher, often as high as 15 times the normal values. There was some relation

between the gravity of the disease and degree of inorg S retention. A diagnostic and prognostic value is attributed to detms of this sort. PETER MASUCCI

Lipoldolysis and tumors. PIETRO ROMONI *Minerva med* 1931, I, 312-8.—The role played by lipoids in the constitution of protoplasm and in cellular multiplication is discussed. R summarizes and reviews his own and the work of others on the relation of lipid metabolism to tumors, and develops the conception that neoplastic proliferation is accompanied by lipoidolysis phenomena. PETER MASUCCI

Renal diabetes. SALVATORE ROMANO *Minerva med* 1931, I, 349-53.—A case of renal and nephritic glucosuria was studied from the standpoint of glucemia during fasting, normal alimentary glucemia, independence of glucosuria from the amt. of carbohydrate ingested and absence of diabetic symptoms. R concludes that renal glucosuria is probably independent of the renal lesion. PETER MASUCCI

Total inorganic calcium and phosphorus in the blood of parathyroidectomized dogs deprived of the large and small intestines with the exception of a portion of the superior duodenum. A BILLI *Boll soc ital biol sper* 6, 6-7(1931).—Inorg Ca in the serum did not diminish. Inorg P increased markedly, reaching a max of 11 mg per 100 cc serum. The ratio Ca/P after the parathyroidectomy diminished appreciably in respect to the value it had before the operation. PETER MASUCCI

The anavenoms. II. Immunization by the anavenoms. MAURICE ARTHUR *J Physiol Path Gen* 28, 773-88(1930) & *C A* 24, 5871.—The anavenoms, prep'd by treating the venoms with formaldehyde, when injected repeatedly under the skin of rabbits, confer a certain degree of immunity against the venoms. Increasing the no of injections has a greater effect in augmenting the immunity than increasing the size of the dose. Cobra anavenom has an immunizing power equiv to the corresponding venom. The antigenic property actually resides in the anavenom and is not the result of a min quantity of venom untransformed by the formaldehyde treatment. III. Anaphylaxis caused by the anavenoms. *Ibid* 800-15.—When rabbits are repeatedly injected with the anavenom of cobra or *Crotalus adamanteus*, a subsequent injection with the corresponding venom will cause anaphylactic shock, exhibited by lowered arterial pressure and increased respiration. This anaphylaxis is no more specific than the anaphylaxis caused by the venoms. Shock can be brought about by injecting cobra venom into rabbits which have been prep'd with the anavenom of *Crotalus* or *Bothrops*. The injection of a dil soln of cobra venom (1/20,000) causes a local, congestive reaction in animals previously injected with anavenom. L. A. MAYNARD

Physicochemistry of the humoral reactions. GAETANO VIALE *Riv sud-americana endocrinol immunol quimioterap* 14, 17-24(1931)(Italian).—Immunity, anaphylaxis, precipitin reaction and other serological reactions are discussed in their relation to pH and the colloidal and chem structure of the compds involved. A. E. MEYER

The bile salts, bilirubin and cholesterol in blood during jaundice. B. VARELA FRENTES, P. RUBINO AND E. AROLO *Semana med* (Buenos Aires) 1931, I, 555-62.—The changes of bilirubin in the blood give the most reliable information about the course of the disease. The normal value of 4 mg glycocholic acid per 100 cc. blood is increased parallel with the bilirubin. A high value in bile salts in blood coexists often with a neg reaction of Ifay in the urine. The blood cholesterol can be higher or lower than normal without evident relation to the disease. A. E. MEYER

Studies on the pathological function of the kidneys in renal disease, especially Bright's disease. I. CAI HOLTEN AND POUL B. REHBERG *Acta Med Scand* 74, 479-518(1931).—An interesting discussion of renal function tests. S. MORGULIS

The presence of proteolytic enzymes in serum. K. YOKOTA *Biochem Z* 232, 58-68(1931).—Human serum may hydrolyze animal fibrin to a small extent, but not human fibrin. Serum from cancerous or laetic patients likewise is unable to hydrolyze human fibrin, nor does it hydrolyze animal fibrin any more than normal serum does. Ligating the pancreatic duct in the rabbit frequently causes a rise in the serum proteases, but the increase is not large. S. MORGULIS

Complement fixation by the interaction of normal serum and bacterial suspensions. A contribution to the study of natural immunity phenomena. T. J. MACKIE AND M. H. FINKELSTEIN *J Hyg* 30, 1-23(1930).—Normal sera from a wide variety of mammals can fix complement with many varieties of bacteria. The property is most marked in man, ox, sheep and horse, but there is much individual variation within a species. The reactions vary in degree with different species and different strains, the strain differences being more important. The reacting principle of the serum is highly thermolabile, being uniformly inactivated at 55° within 30 min. Complete inactivation often occurs at 50° to 52° and partial inactivation at 40°. The complement-fixing substance is "absorbed" by bacteria at 0°, which thus become sensitized. This substance can

be absorbed by other substances than bacteria but not so completely. As a rule the reacting power of bacterial suspensions is destroyed at 55°, but this is influenced by different conditions as the type of medium on which the cultures are grown. The reacting power of the serum is fully developed in young animals. These natural antibody-like principles differ from immune antibodies in thermostability. JOHN T. MYERS

Transmission of maternal immunity. J. H. MASOV, T. DALLING AND W. S. GORDON. *J Path Bact* 33, 783-83(1930).—In sheep, cattle and horses, antibodies do not pass through the placenta into the fetal circulation. They do appear in the colostrum and are thus absorbed by the offspring during the first few days after birth. JOHN T. MYERS

The absence of trypanocidal substances in human cerebrospinal fluid. P. REGENDANZ. *Zentr Bakt Parasitenk*, I Abt., 120, 89-91(1931) JOHN T. MYERS

The antihemolytic properties of human serum. M. C. RUBINO-MONTAVIDEO. *Zentr Bakt Parasitenk*, I Abt., 120, 145-50(1931).—Spontaneous inactivation (3 to 5 days in the ice box) destroys complement, but not the anticomplementary properties, which on the contrary increase. Heating to 56° destroys both properties. J. T. M.

The dissociation of the diphtheria anatoxin-antitoxin complex and the recovery of the anatoxin. G. RAMON, R. LECROUX AND M. SCHOFF. *Compt rend* 192, 512-4(1931). *C A* 24, 2173.—Diphtheria anatoxin behaves like toxin in its power to form a complex with diphtheria antitoxin, and in the possibility of dissociation of this complex in the same manner as that of the toxin-antitoxin complex. T. H. RIDER

The ultimate fate of arrested grafts of epithelial tumors. F. VILS AND A. DE COLOV. *Compt rend* 192, 541-3(1931).—Epithelial tumors grafted in mice are often inhibited and apparently disappear by a complete lysis. Mice in which a grafted tumor had disappeared were painted with tar at another site, after as long as 120 days, causing, in some cases, a new tumor at the site of the original graft. Others showed growths at the site of the painting. Only 2 animals showed both. T. H. RIDER

The metabolism of neoplasm tissue. ANNA GOLDFEDER. *Med Dermatol* 13, 81-89(1931)(in Russian, English Abstract).—Neoplasm tissue always contains less glycogen (1) than normal tissue. It favors the growth of neoplasm. Notwithstanding the presence of lactic acid, neoplasm tissue has a *pn* of 7.5-7.7, because of the production of NH_4 . T. H. RIDER

Spasmophilic forms of infantile enuresis. G. MACCIOTTA. *Pediatrics russa* 38, 1145-57(1930).—The investigations were carried out on 24 children with enuresis; the children showed no other symptoms which could be regarded as the cause for the enuresis. In 17 of the cases M found a decrease of blood Ca, a slight diminution in the H-ion concn of the blood and a more or less marked increase of the galvanic excitability, with inversion or equalization of the anodic formula. M concludes from his findings that in these cases the enuresis is a manifestation of an existing spasmophilia. The therapy with ultra-violet rays, alone or in combination with parathyroid hormone, gave good results. The best results were obtained by the combined treatment with ultra-violet rays and vitamin B. Numerous references are given. G. SCHWACH

The possibility of the occurrence of anaphylactic phenomena in the serum therapy by the oral route. LUIGI AURICCHIO. *Pediatrics russa* 39, 289-99(1931).—Five cases of alimentary serum anaphylaxis are reported. In 3 of these cases sensitization was produced by parenteral administration of horse serum, slight, transitory disturbances occurred after oral administration of horse serum preps. In the 2 other cases sensitization was caused by oral administration of such preps and very severe anaphylactic reactions, which led to death in 1 case, were observed after injection of serum. Expts. carried out in guinea pigs showed that specific sensitization can be produced by prolonged gastric administration of horse serum. Such animals exhibited only slight and transitory symptoms when they were injected with cattle serum, while with horse serum anaphylactic symptoms appeared slowly until the animals died on the 3rd or 4th day after the injection. Gastric administration of horse serum was without influence on guinea pigs sensitized by injection of horse serum. G. SCHWACH

The electric conductivity of the cerebrospinal fluid in normal pathological conditions. P. GORINI. *Pediatrics russa* 39, 300-8(1931).—G detd the elec. cond. of the cerebrospinal fluid in 73 infants. In the healthy infants *K* was fairly const (168-183 at 35°). Normal values were also found in acute and chronic diseases and in cases with pathol alterations of the nervous system excepting meningitis, in which condition low values for *K* were found (min 143). G. SCHWACH

Acute insufficiency of the adrenal glands. AUGUSTO ANTONINI. *Rass clin terap sci affm* 29, 329-50(1930).—A review. Four pages of references are given. G. S.

The nature of variation of agglutination titer caused by insulin. M. NOZAKI.

Kelso J Med 1, 248 87(1930) —Insulin injected intravenously into rabbits immunized with typhoid bacilli causes a great but temporary increase in its agglutination titer. This increase is considered as not being due to the insulin itself but rather dependent upon hormonal action. A J VORWALD

Serology of syphilis. IV. A more sensitive antigen for use in the Wassermann reaction. HARRY LAGLE *J Exptl Med* 53, 605-14(1931), cf *C A* 25, 139-40 —The discovery that there are many substances with the sensitizing properties hitherto believed peculiar to cholesterol and its derivs and that sensitizer can be added to antigen in very large quantities, many times those currently used, and yet continue to increase its complement fixing efficiency with no danger of giving falsely pos tests, has made possible the prepn of an antigen much more sensitive than any now available for use in the Wassermann reaction. This consists of the EtOH ext of beef heart muscle to which are added 0.8% cholesterol and 0.6% sitosterol, the excess of which crystallizes out on cooling and is dissolved before using. Details of the prepn are given. There is reason to believe that this antigen possesses almost the max sensitivity obtainable. The method of prepn insures its being almost cold with antigen-lipoids, and more sensitizer could not be added without increasing the turbidity of its soln in saline to a point where it would interfere with the reading of hemolysis. V The cause of the greater sensitivity of the ice-box Wassermann, the zone phenomenon in complement fixation. *Ibid* 615-22 —Serum, in concns greater than 1:25, causes a marked inhibition of complement fixation in general and of the Wassermann reaction in particular. The serum protein is probably adsorbed by the colloidal dispersed lipid reagin complexes, forming a protective film which prevents the fixation (adsorption) of complement. This inhibition explains the zone phenomenon in complement fixation: a weakly pos serum may give a completely pos reaction in, e.g., 1:5 diln, and yet, because of this serum inhibition, may appear completely neg when tested as whole serum. The greater sensitivity of the ice-box test is due to the fact that the serum inhibition just described is less marked at lower temp., to the prolonged incubation time, making for greater sp fixation, to a more marked non-sp destruction of complement by antigen and a spontaneous deterioration in the longer ice box test. Because of the inhibition by serum protein in high concn, a quant Wassermann technic involving the use of graded quantities of serum is worthless when carried out at 37°. Even the ice-box test, which is less susceptible to this inhibiting effect, will yield a pos reaction with whole serum only when the circulating reagin exceeds a surprisingly high threshold (6-10 times the quantity which could be detected in dil serum). C J WEST

Inflammation. VII. Fixation of bacteria and of particulate matter at the site of inflammation. VALY MINKIN *J Exptl Med* 53, 647-60(1931), cf *C. A.* 25, 1570 —India ink or graphite particles injected into an area of inflammation fail to disseminate to the tributary lymph nodes. When injected into a normal peritoneal cavity they rapidly appear in the retrosternal lymph nodes. When injected into an inflamed peritoneal cavity, they are fixed *in situ* and fail to reach the regional lymph nodes. Graphite particles injected into the circulating blood stream enter an inflamed area both as free particles as a result of increased capillary permeability and also as phagocytized material within leucocytes. *B. prodigiosus* injected into inflamed tissue is fixed at the site of inflammation and fails to disseminate to the regional lymph nodes as readily as when injected into normal tissue. Injected at the periphery of an inflamed area it does not readily penetrate into the site of inflammation. Injected intravenously, it rapidly enters an inflamed area. C J WEST

Congenital protein hypersensitiveness in two generations. BRET RATNER AND HELEN L. GRUEHL *J Exptl Med* 53, 677-86(1931) —Hypersensitivity actively induced *in utero* is shown to persist for a longer period than passive sensitization (cf *C A* 21, 449). The degree of hypersensitivity, its duration and its transmissibility appear to be influenced by the time elapsing between the original injection of the parent and parturition. A pregnant guinea pig receiving a parenteral injection of antigen 2-4 days prior to parturition transmits a state of hypersensitivity to 2 succeeding generations. The sensitization of the F¹ generation is due to the passage of antigen. The sensitization of the F² generation is due to the passage of antibodies formed in the F¹ generation. This prevents any further transfer of the hypersensitive state. Though hypersensitivity occurs in 2 successive generations, the phenomenon is congenital and not hereditary. C J WEST

Lipoid metabolism. I. Lipoid metabolism in rabbits with fever. HIDEAKU HAMANO *Proc Imp Acad Tokyo* 7, 80-1(1931) —The total fatty acid of the whole blood and plasma is decreased in fever, the max being reached in 24 hrs. The whole blood total fatty acids return to normal in 60 hrs, the plasma in a shorter time. Choles-

terol and lecithin increase both in whole blood and in plasma in the beginning of the fever; the cholesterol reaches a max. in 24 hrs., the lecithin in 6 hrs.; after this time both decrease and fall below the normal value. The serum protein decreases with the fever and becomes normal in 15-24 hrs.

C. J. WEST

Daily variations in sugar content of blood and urine in normal and in diabetic persons. J. MÖLLERSTRÖM. *Svenska Läkarförbundet Handlingsnr* 56, 211-95 (1930); *J. Am. Med. Assoc.* 96, 2271. —About 12,000 blood-sugar tests and 7000 urine tests were made on 37 patients. In glucosuria the amt. of sugar excreted varied from day to day even on a const. diet. The excretion was sometimes periodic; thus in diabetics a min. was often noted at midday, even with frequent meals. Excretion of acetone bodies in grave diabetics was also periodic. At certain times of the day, both in normal and diabetic persons, postalbuminuric glucosuria appeared more rapidly or did not appear. During this stage there was a tendency to spontaneous fall of the blood sugar, sometimes so marked that in spite of ingestion of food the blood sugar was still below the starvation level.

R. C. WILLSON

KINDYER, ROBERT A. The Clinical Interpretation of Blood Examinations. Philadelphia: Lea & Febiger. 623 pp. \$6.50. Reviewed in *J. Am. Med. Assoc.* 96, 1572 (1931).

THOMSON, OLUF. Antigens in the Light of Recent Research. London: Heinemann, 14s. net.

N-PHARMACOLOGY

A. V. RICHARDS

Secondary and toxic effects of arsenobenzene preparations. F.M.L. BREZOVSKY. *Wiener klin. Wochenschr.* 43, (28): 7 (1931). —The preps. studied and the % of cases in which secondary effects were negligible were: *acetoarsophenine* 37, *arsenal* (arsenolite) 10, *arsenical* (phos.) 41, *arsenobenzene* (Bism.) 6, *arsenophenine* 40, *myoarsophenine* 4, *arsenical liquid* 19, *arsen* 10.

D. B. DILL

Metal salt therapy. Concluding experiments with rabbit tuberculosis. L. E. WALSH. *Z. Tuberk.* 53, 292-9 (1931). d. C. A. 24, 435. —On the basis of extensive observations with 42 metals and metal salts the author concludes that Ca and Mn possess a sp. healing value in rabbits infected with tubercle bacilli while only a slight action was noted with Co, Ba, Al, La, Mo, Pt, Ni and Sm, the remaining 32 being without effect.

H. J. COOPER

Metallic electrocoagula in the treatment of syphilis. ORLANDO RANGEL. *Ed. esp. brasil. pharm.* 11, 3-22 (1931). —Some advantages are given of the use of electrocoagula (Hg and Bi Hg) over arsenicals in the treatment of syphilis.

C. A. A.

Clinical test of a new compound in the treatment of arthritis. D. E. MARKSON. *Am. Med.* 37, 116 (1931). —*Monodichrophen* was found advantageous in the treatment of arthritis.

RACHEL BROWN

Field experiences in the use of calcium gluconate. J. ALLEN ROGERS. *J. Am. Vet. Med. Assoc.* 31, 555-6 (1931). —A report of the successful treatment of milk fever in cows by the injection of a 10% soln. of Ca gluconate.

RACHEL BROWN

The curative value of certain bromine compounds. An experimental study. SOMMER S. GREENEARTH WITH ANNA M. RILEY. *Am. J. Syphilis* 15, 59-71 (1931). —The curative doses of 12 commercial Br preps. were detd. by the lymph-gland method after intramuscular injections in rabbits. There was variation in the curative dose from 2 mg. per kg. ("bromthional") to more than 25 mg. per kg. ("bromidol" and "iodobromthional").

B. C. BRUNSTETTER

Intravascular phagocytosis of erythrocytes in Necturus following prolonged immersion in lead acetate. A. B. DAWSON. *Anat. Record* 45, 345-56 (1931). —Extensive phagocytosis of erythrocytes occurs in the circulating blood of Necturus following prolonged immersion (15 days or more) in a soln. of Pb(C₂H₃O₂)₂. Phys. changes in the erythrocyte, involving loss of typical form and the acquisition of the quality of stickiness (to glass) appear to favor phagocytosis. The destruction of erythrocytes was accompanied by striking changes in the color of the bile.

B. C. BRUNSTETTER

The effect of administration of thyroid substance, of potassium iodide, and of excision of the greater portion of the thyroid gland on the epidermis of the guinea pig. LEO LOEZ, F. L. HAVEN, I. T. GENTHER AND HILDA FRIEDMAN. *Anat. Record* 46, 65-80 (1930). —Administration of thyroid substances or of KI daily to guinea pigs did not cause an increase in the proliferative activity of the epidermis or an increase in the number of epidermal cells. The results suggest that both multiple as well as single doses of these substances may cause a lowering of the proliferative activity and

number of cells. Following a not quite complete thyroidectomy no lowering of the proliferative activity in the epidermis nor of the number of epidermal cells took place. The results are discussed with reference to their bearing on the problem of the significance of the internal secretions on the origin of cancer. B C BRUNSTETTER

The chromatophores of cephalopods. III. The effect of some poisons in vitro. E SERENI. *Z. vergleich. Physiol.* 12, 329-503 (1930).—The action of the following substances on the nerve, the muscle and the zone between nerve and muscle, of nerve-muscle preps was investigated: aconitine, atropine, acetylcholine, betaine, caffeine, cocaine, eholine, conine, ephedrine, guanidine, lobeline, phenol, pilocarpine, strychnine, strophanthine, tyramine, veratrine and quinine. It was demonstrated that the zone between nerve and muscle is differentiated from nerve or muscle not only pharmacologically but physiologically. One section of the nerve, the intermediate zone can be excited only by specific chemical stimuli. It is probable that the conduction of impulses from nerve to muscle occurs normally through the intermediate zone by means of a chem. mechanism. Tyramine and betaine both exert a direct effect on the chromatophores, the first by increasing the excitability of the muscle (thereby resulting in the appearance of pigment clouds) and the second by stimulating the zone between nerve and muscle (which causes a contraction). As both substances occur in cephalopods, it is assumed that they control at any instant the condition of normal chromatophores by their action on the nerve centers. Pharmacological analysis confirms the idea that the diffuse expansion of pigment immediately following the application of an elec. stimulus to normal cephalopod skin is due to a stimulation of nerve, and not of muscle, fibers. B C BRUNSTETTER

Notable cases of poisoning. R FRIML. *Pharm. Monatshefte* 12, 61-2 (1931).—An address enumerating certain of the more notable cases of intoxication, either by intent or accident, from ingestion of preps. contg. Hg, Cu, As, Ba, Pb, Th, Zn, I, HCl, CO₂, EtOH, CH₃OH, acetylsalicylic and salicylic acids, pyramidone, veronal, dial, luminal, opium alkaloids, quinine, atropine, castor oil seed, mushrooms, etc. W O E

The chemotherapeutic action of quinine (and its derivatives) in vitro upon *Proteosoma praecox* (avian malaria). W BORCHARDT. *Arch. Schiffs Tropen Hyg.* 34, 380-4 (1930).—Quinine, hydroquinine-HCl and plasmochin incubated with the blood of infected animals *in vitro* have a definite parasitocidal action, the lethal doses being 0.5, 4, and 0.2 mg, resp., per cc of blood (5 hrs. at 37°). These expts. indicate that the therapeutic effect of quinine derivs. in malaria is due to a direct action upon the parasite *in vivo*. H EAGLE

Experimental studies on the pharmacology of salyrgan. IV. Studies on the chemistry of salyrgan. KNUT O MÖLLER. *Arch. expul. Path. Pharmacol.* 153, 109-19 (1930), cf. C A 24, 5067.—As shown by potentiometric detns. against a Hg electrode, salyrgan produces practically no Hg ions at room temp. The addn. of acid causes a voluminous ppt. of "salyrganic acid." This is only slightly sol. in H₂O (1:58000), and somewhat more sol. in acid (1:4600 in 2 N H₂SO₄). In acid soln. the Hg ion is formed. Salyrganic acid is sol. in NaCl soln., probably as a result of the formation of a complex salt. The Hg ion concn. decreases with increasing NaCl concn. The acid is also sol. in KCl, NH₄Cl, CaCl₂, MgCl₂, BaCl₂, NaI and NaBr, but only very slightly sol. in Na₂SO₄, NaNO₃, ZnCl₂, FeCl₃ and HgCl₂. Salyrgan is excreted in alk. urine as such, passing through the body unchanged. It is excreted in acid urine as the NaCl salyrganic acid complex. H EAGLE

The effect of Germanin upon the blood-coagulation system with particular reference to its prophylactic and therapeutic use in thrombosis. BERNHARD STILBER AND KONRAD LANG. *Arch. expul. Path. Pharmacol.* 154, 41-9 (1930).—Intravenous injection of Germanin causes a slight shift in the reaction of the blood toward the alk. side and a relative increase in the albumin fraction of the serum protein. There is a decrease in the thrombocyte count and an increase in their charge which inhibits their agglutination. Germanin has been successfully used in thrombosis and thrombophlebitis. H EAGLE

The effect of the intravenous infusion of sugar upon the coagulation system, with particular reference to the pathogenesis of thrombosis. BERNHARD STILBER AND KONRAD LANG. *Arch. expul. Path. Pharmacol.* 154, 50-8 (1930).—The intravenous infusion of glucose causes an increase in blood fibrinogen and globulin, an acceleration of coagulation, an increase of thrombocytes and a decrease in their elec. charge. The role of these changes in causing thrombosis is discussed. H EAGLE

Comparative studies on the effects of phenol and some phenol derivatives upon muscle. TAKATOSHI HASEGAWA. *Arch. expul. Path. Pharmacol.* 154, 103-14 (1930).—The narcotic effect of various phenol derivs. upon surviving muscle preps. can be

readily measured by the time required for complete passivity of the isolated frog gastrocnemius to maximal stimulation. The effect of phenol, *p*- or *o*-cresol, and *m*-cresol are in the ratio of 1:2.5:1.75, resp. In contrast, diphenols excite the muscle, and triphenols are only slightly active, but their ethers are pure narcotics. Thymol, β naphthol, mono- and poly nitrophenols have no narcotic effect and cause an intense contraction which develops into death rigor, possibly due to protein coagulation. Amino- and sulfo-phenol are ineffective, while benzene causes both narcosis and reversible contraction. H. EAGLE

The localization of the waking effect of calcium in magnesium narcosis. RUDOLF SCHÖEN AND SIGGRIED KÖRFFEN. *Arch exp'l Path Pharmacol* 154, 115-27 (1930). — Mg narcosis is due to paralysis of the proprioceptive and labyrinthine reflexes, the waking effect of Ca persists as long as the reflex centers are intact, and persists even if the thalamus, corpus striatum and cerebrum are removed from the exptl rabbit. The rigidity following deceleration is removed by Mg, but restored by Ca. The waking effect of Ca upon the Mg narcosis is therefore not dependent upon an intact corpus striatum, and is not due to a quieting effect upon the brain stem. H. EAGLE

Direct and antagonistic effects of magnesium, calcium and potassium chlorides and sodium oxalate upon respiration. SIGGRIED KÖRFFEN. *Arch exp'l Path Pharmacol* 154, 129-42 (1930). — Given intravenously, CaCl_2 , MgCl_2 and KCl all increase respiration, given subcutaneously, the first two paralyze respiration, the third is ineffective. The respiratory paralysis caused by MgCl_2 is instantaneously stopped by intravenous CaCl_2 , gradually overcome by subcutaneous CaCl_2 , and increased by $\text{Na}_2\text{C}_2\text{O}_4$. Respiratory paralysis due to CaCl_2 is removed by the intravenous injection of MgCl_2 . Neither KCl , CaCl_2 nor MgCl_2 affects the respiratory paralysis caused by urethan or dial. H. EAGLE

The pharmacology of gluconic acid: a contribution to the effect of free acids in the body. I. SIGMUND LIPPMANN. *Arch exp'l Path Pharmacol* 154, 143-60 (1930). — *Bact. gluconicum* was grown in media contg glucose, the gluconic acid formed was repeatedly recrystd. as the Ca salt, and the Ca was removed with H_2SO_4 . Various bacteria oxidize the free gluconic acid to ketogluconic acid. The toxicity of the free acid upon intravenous injection depends upon its concn., and is less if it is given in a soln. contg NaCl . Org. acids in general differ greatly in their toxicity, even when differences in $[\text{H}]^+$ and in molar concn. are eliminated. One important factor in this is that their very rapid oxidation in the body makes for hypotonicity, partially explaining the protective action of NaCl . II. *Ibid* 161-74. — Org. acids vary in their ability to form hematin from defibrinated blood, moreover, the activity of each acid depends upon the species of animal from which the blood was taken. Although there is a certain correlation between the effective H^+ ion concn. and molarity, roughly, $K = m[\text{H}^+]^2$, these are not the only two factors: the ability of the acid to penetrate into the red cell is very important. II, instead of using the intact cell, one uses a soln. of hemoglobin, the activity of the acid depends upon the square of $[\text{H}]^+$, and is entirely independent of the molarity of the acid. The very slight activity of gluconic acid upon the intact cell is due to the fact that it is least sol. in lipoids, and therefore penetrates least into the cell. III. *Ibid* 175-92. — The intravenous injection of org. acids causes an increase in minute respiratory vol. and frequency, even in the respiratory depression due to morphine. Most of the acids, exclusive of gluconic acid, cause a temporary fall in blood pressure. Gluconic acid has no significant effect upon the Ca metabolism of rabbits and dogs. In the former the excretion of Ca in the urine and feces is increased, in the latter, urine Ca is increased and feces Ca decreased. The parenteral administration of gluconic acid causes an increase in serum Ca, i. e., a mobilization of Ca, which explains the fact that it decreases the toxicity of citric acid. Only a fraction of gluconic acid given by mouth is absorbed as such, because of decomposition by yeast and bacteria in the intestine. H. EAGLE

The antagonism of a adrenaline and insulin in the regulation of blood sugar (experiments upon adrenalectomized dogs). F. MEYTHALER AND R. CARIO. *Arch exp'l Path Pharmacol* 154, 193-202 (1930). — Injections of concd glucose solns into the pancreaticoduodenal artery in adrenalectomized dogs causes a fall instead of a rise in blood sugar, just as it does in normal dogs. Injection into peripheral vessels, however, causes a temporary marked hyperglucemia. This confirms the results of Grafe and Meythaler, who view the blood sugar as the adequate hormone for the secretion of insulin, quite apart from any effect of adrenaline. There is an antagonistic action between the two in the regulation of blood sugar. H. EAGLE

The effect of nicotine, pilocarpine and veratrine upon the isolated submaxillary salivary gland of the dog. O. NIKOLAEV. *Arch exp'l Path Pharmacol* 154, 202-10

(1930)—Perfusion of the 3 drugs through the isolated gland has the same effect upon salivary secretion and blood vessels as in the intact animal. H. EAGLE

Oxidation of benzene in the isolated livers of warm- and cold-blooded animals. CHERNIKOV, IDA D. GADASKIN AND I. I. GUREVICH *Arch. expil. Path. Pharmacol.* 154, 222-7 (1930).—The isolated livers of cats, rabbits and frogs oxidize benzene to phenol. In benzene poisoning of the liver of warm-blooded animals, phenol is found in both free and bound forms, after benzene poisoning of frogs, phenol is found in the cloacal contents. The liver may not be the only organ concerned in this oxidation. H. EAGLE

The effect of snake poisons upon nerve tissue. ARTHUR WEIL *Arch. expil. Path. Pharmacol.* 154, 228-38 (1930), cf. *C. A.* 24, 3267.—The poison of *Naja nigricollis* acts like sapotoxin, that of *Bothrops atrox* is similar to Na taurocholate. Both affect the medullary sheath, without affecting the affinity of the axis-cylinder for Ag stains. A typical Marchi reaction (reduction of OsO_4) is given by the fibers affected by sapotoxin and both snake poisons, but not by fibers injured with Na taurocholate. The adsorption of sapotoxin or of the snake toxins by the spinal cord follows Arrhenius' formula for the adsorption of amboceptor by red cells rather closely, as well as Schütz' rule. H. EAGLE

The disappearance of the hypertensive action of the hypophysis in the body. WERNER HARTMANN *Arch. expil. Path. Pharmacol.* 154, 255-62 (1930).—The vasopressor action of posterior pituitary exts. is less marked if it is injected into a peripheral artery than if it is given intravenously. Similarly, passage through the portal capillaries removes some of its activity. H. EAGLE

Experimental studies in the pharmacology of salyrgan. V. The behavior of mercury in the body after the administration of salyrgan. Continued studies on salyrgan nephritis in rabbits. KNUD O. MØLLER *Arch. expil. Path. Pharmacol.* 154, 263-70 (1930).—After the intravenous injection of salyrgan into dogs and rabbits, there is a very rapid excretion of Hg in the kidneys, far more rapid than in man (40% in one hour, 70-80% in 2 hrs). The urine Hg concn may be as high as 90 mg. % in the dog and 370 mg. % in the rabbit. The development of salyrgan nephritis in the rabbit does not depend upon the reaction of the urine, moreover, the course of Hg excretion is the same in acid as in alk. urine. Only 3.5% of the Hg is excreted by the liver in 24 hrs in the dog, and only 1% is excreted through the walls of the intestine. After repeated injections of salyrgan into dogs only minimal quantities can be found in the organs, chiefly in the kidneys and liver. H. EAGLE

The sugar economy of the liver. I. The effect of adrenaline and insulin upon the sugar output and glycogen content of the liver. HANS MOLLER AND LEO POLLAK *Arch. expil. Path. Pharmacol.* 154, 280-300 (1930).—The difference between the sugar contents of the hepatic vein and of the ear vein is increased after adrenaline administration, simultaneously, there is a decrease in the liver glycogen. There is no quant. relationship, however, between these two effects. After insulin, the sugar content of both the hepatic and ear veins falls within 10 min., but the carbohydrate content of the liver remains either approx. const., or shows a slight decrease. Adrenaline therefore increases the glucose output of the liver, while insulin suppresses it. H. EAGLE

Clinical and experimental contributions to the pharmacology of harmine. G. MARINESCO, A. KREINDLER AND A. SCHEIM *Arch. expil. Path. Pharmacol.* 154, 301-16 (1930).—Harmine causes clinical improvement in cases of Parkinson syndrome. In normal subjects and in post-encephalitic parkinsonism it causes an increased excitability of the vestibular app. There is a fall in blood pressure and bradycardia, and the vascular reflexes return toward normal. The hyperexcitability of the parasympathetic system and the lowered excitability of the sympathetic system characteristic of parkinsonism are also corrected by harmine. It has a vagal effect upon isolated frog heart, and reduces muscle and nerve chronaxia in the isolated frog nerve muscle preps. In the patient it has a therapeutic effect upon the pathological chronaxia of nerves and muscles. H. EAGLE

The cause of creatinuria. I. Acidosis and creatinuria. OTTO RIESSER AND CARLO BRENTANO *Arch. expil. Path. Pharmacol.* 155, 1-20 (1930).—Despite the fact that morphine, urethan and ether narcosis induce both creatinuria and acidosis, the two are not necessarily correlated. The degree of creatine formation and excretion is not proportional to the degree of acidosis and alkalosis induced by the administration of alkali may also cause creatinuria. H. EAGLE

Pharmacology and pharmacodynamics of salves and medicaments incorporated in salves. VI. The pharmacology and pharmacodynamics of unguentum-hydrargyri praecipitati albi, Ph. g. CARL MONCORPS *Arch. expil. Path. Pharmacol.* 155, 51-69

ates the poisoning of the frog muscle independently of pH . There is no relation between the toxic effects of hydroquinone and its effects upon tissue respiration. Muscles of fatigued animals or of animals treated with thyroxine are more susceptible to poisoning by hydroquinone, lactic, succinic and pyruvic acids depress this susceptibility, and cane sugar accentuates it. H EAGLE

The clinical significance of the initial insulin hyperglucemia. MAX BÜRGER *Klin Wochschr* 9, 104-8(1930)—In 22 normal fasting human beings, the intravenous injection of insulin caused an initial hyperglucemia, the max rise averaging 22% of the original blood sugar, this max was reached in 10 min. In cirrhosis, this initial hyperglucemia is not observed, in jaundice, it is significantly decreased. Both an intact liver parenchyma and a liver store of glycogen are considered necessary for the phenomenon, this explains its absence in Basedow's disease, and its presence in myxedema. In diabetes under proper treatment hyperglucemia is also observed, but not in diabetic coma. H EAGLE

The vasoconstrictor action of carbonic acid and its significance for the circulation. KL. GOLLWITZER-MEIER AND HANS BOHN *Klin Wochschr* 9, 872-5(1930)—The inspiration of air with a high CO_2 content by dogs causes vasoconstriction, demonstrable even in a vein being perfused with a normosal-serum mixt, quite apart from the circulation. This constriction is conditioned by a vasomotor reflex, and is not observed in the perfused vein if its nerve supply is eliminated. H EAGLE

The retention of uroselectan in the human body. II. The distribution of uroselectan in the animal body. W. TORNÉ AND C. DAMM *Klin Wochschr* 9, 1719-20(1930), cf C A 25, 1582—Thirty five min after the intravenous injection of 10-30 cc of 40% uroselectan into rabbits the animals were killed and their organs examined. Uroselectan was found to be stored chiefly in the liver, kidneys, skin, blood and urine. A similar distribution was found in dogs. The kidney content is ascribed to urine, that of the liver to bile. Damage of either the liver or the kidneys leads to storage and excretion in the other organs. Widespread kidney damage is a contraindication to the administration of uroselectan. H EAGLE

The action of vasopressin and oxytocin upon intestinal peristalsis and the treatment of ileus paralyticus with vasopressin. A. WL. ELMER, L. PTASZK AND M. SCHIEPS *Klin Wochschr* 9, 1763-8(1930)—Of the hormones present in posterior pituitary exts., vasopressin stimulates, and oxytocin inhibits, intestinal contractions. The former is more active than pituitrin, physostigmine, pilocarpine, atropine, etc., and its therapeutic use in paralytic ileus is indicated. H EAGLE

The pharmacological significance of the sterols. HANS SEEL *Klin Wochschr* 10, 241-4(1931)—A review. H EAGLE

The properties and effects of a new class of organic bromine compounds, SIEGWART HERMANN AND MARIE FREUND *Klin Wochschr* 10, 230-2(1931)—If Br_2 is combined with unsatd fatty acids (e.g., oleic acid), and part of the Br_2 is removed, one attains a new type of org. combination of Br_2 , pharmacologically many times more active than the fully brominated compd. A quantitatively similar relationship was observed with Cl_2 . The pharmacol. effect of Br_2 compds usually depends upon the type of compd rather than the amt of Br_2 . Multibrol (Na monobromoleate) has no effect; it does not affect blood pressure, respiratory vol or frequency, and does not cause bromism. H EAGLE

The effect of caffeine in hypoglucemic conditions. GRIEFF AND HAPPE *Klin Wochschr* 10, 263(1931)—Contrary to the findings of Popper and Jahoda, the symptoms following an excess of insulin depend upon the fall of blood sugar. Caffeine alleviates these symptoms, without affecting the blood sugar. There is no reason for assuming that insulin contains two distinct substances, one with a hypoglucemic and the other with a toxic action. Such symptoms as urticaria, and possibly convulsions, may be due to protein impurities. H EAGLE

The circulatory effect of ephedrine. F. HILDEBRANDT AND H. MCCGE *Klin Wochschr* 10, 291-4(1931)—The circulatory effects of ephedrine are due not to an increase in cardiac activity, but to a direct action upon the vessel walls. Ephedrine may have a toxic action upon the heart. H EAGLE

The appearance of indican and aromatic substances in the blood in severe cardiac decompensation and morphine poisoning. J. E. W. BROCHER *Klin Wochschr* 10, 294-7(1931)—In a case of morphine poisoning, coupled with myocardial insufficiency, there was marked retention of urea, uric acid and indican in the blood, with a prompt return of the abnormal blood findings to normal when diuresis was established. H E

Günsslen's liver extract for injection. Further evidence against liver resistance

and in favor of the susceptibility of cord degeneration to treatment. VIKTOR SCHILLING *Klin Wochschr* 10, 301-3(1931). H. EAGLE

The effect of atropine upon glutemia in human beings. D. DANIFLOPOLU *Klin Wochschr* 10, 311-3(1931) — Contrary to current belief, stimulation of the parasympathetic system increases the blood sugar. H. EAGLE

The treatment of pernicious anemia with a stomach extract which can be injected. W. I. DERLE, H. KRAFFT and M. GÄSSLEN *Klin Wochschr* 10, 313(1931) — Good results are reported in this preliminary report. H. EAGLE

A new mechanism of the insulin effect. H. SCHWARTZ and M. TAUBENHAUS *Klin Wochschr* 10, 313-4(1931) — A preliminary report. H. EAGLE

Thyroxine and percutaneous sensitization. M. LUCAS *Klin Wochschr* 10, 314(1931) — Thyroxine accentuates the allergic reaction due to percutaneous sensitization, and therefore has a peripheral action. H. EAGLE

The specific treatment of gonorrhea with soluble gonotoxin. W. WOLFFENSTEIN *Klin Wochschr* 10, 354-6(1931). H. EAGLE

Experimental studies on the effect of the inspiration of small quantities of [benzene] and [benzene] upon the respiratory organs and the whole body. L. TR. LARIONOV and N. V. LAZAREV *Klin Wochschr* 10, 356-7(1931) — The authors take issue with the conclusions of Schmidtman (C. A. 25, 1006). H. EAGLE

The insulin reaction of the leucocytes of the blood. W. STOCKINGER and K. KOBEL *Klin Wochschr* 10, 389-92(1931) — Within 2 hrs following the administration of insulin, there is an abs. rise in the no. of leucocytes, reaching a max. at the end of the 2nd hr. Simultaneously, or a little later, there is a neutrophilic leucocytosis. There is no correlation between the hypoglycemia and these changes in the blood picture. H. EAGLE

The treatment of lymphogranulomatosis inguinalis, particularly with a copper preparation. W. FRIED and J. WIESE *Klin Wochschr* 10, 401-6(1931) — Although hb. preps are effective, they work very slowly and irregularly. The authors suggest the use of cuprosol H (a proprietary prepn) as an adjunct in therapy, particularly effective if given with Stibeny. H. EAGLE

Ephedrine in local anesthesia. M. KOCHIMANN *Klin Wochschr* 10, 408(1931) — Ephedrine is valuable as an adjunct to local anesthetics, but cannot replace adrenaline. H. EAGLE

Further studies in the demonstration of Germanin (Bayer 205) in the animal body. HEINZ ZEISS and KENIA UTKINA-LYUBOVITSOVA *Z. Immunitäts* 68, 170-81(1930) — Following the intravenous injection of Germanin, the highest concns. are found in the serum; the drug is also stored in the kidneys, lungs, spleen and liver. H. EAGLE

Colloid-chemical combination reactions of Germanin and protein complexes. II. The effect of Bayer 205 upon the protein components of the hemolytic system. VASSILI A. NOVOSSELSKI *Z. Immunitäts* 68, 323-41(1930) — Germanin has no effect upon the hemolytic properties of normal or sensitized sheep cells and does not decrease the hemolytic titer of an amboceptor serum. It does, however, depress the hemolytic properties of complement. H. EAGLE

The residual quantities of arsenic remaining in animal organs after the intake of neoarsphenamine. L. REMY *Z. Immunitäts* 68, 395-402(1930) — Following the injection of neoarsphenamine into rabbits, considerable quantities of As are stored in the organs, particularly in the spleen, lymph glands, liver, bone marrow and brain for sometime. R. opens the question as to whether the excess As is ever completely excreted. H. EAGLE

3,4-Dihydroxyephedrine and 3,4-dihydroxynorephedrine. RAYMOND HAMET *Compt. rend.* 192, 300-2(1931), cf C. A. 25, 1590 — Both 3,4-dihydroxyephedrine (I) and 3,4-dihydroxynorephedrine (II) when injected produce a hypertension which, like that produced by adrenaline (III), is increased by cocaine. The effect of ephedrine (IV) is decreased by cocaine. Large doses of IV are hypotensive but large doses of I and II are hypertensive like III. While IV in small doses has but a slight transient effect on intestines *in vitro*, I and II, like III, have a strong and lasting inhibiting effect. Conclusion — The addn. of a Me group to the β C of III or noradrenaline does not greatly modify the physiol. action of these substances, while substitution in the nucleus of ephedrine of 2 phenolic OH groups in the 3,4-position profoundly changes the action of this compd. L. E. GILSON

Lectures in materia medica. W. A. DEWEY *J. Am. Inst. Homoeopathy* 23, 1224-34(1930), 24, 186-98(1931), cf C. A. 24, 5377 — A discussion of the physiol. action and therapeutic use of *Siramonium*, *Solanum dulcamara*, tobacco, *Capsicum anuum* and *Nux vomica*. JOSEPH S. HEPBURN

The influence of the intravenous injection of urea on the exchange of substances between the blood and the tissues. H. DONOVAN AND D. BRENNER *Brit J Exptl Path* 11, 419-38(1930).—Five g. urea was injected intravenously into healthy individuals and patients with various diseases, and the urea, NaCl and H_2O contents of the blood and urine were detd. at intervals for a period of 2 hrs after the injection. By the time the urea has mixed with the blood (3 min.) 80% has left the blood stream. It is suggested that the urea leaves the blood by a process of simple diffusion, passing into the water of the cells as well as into the free water, and that the process is complete in 15 min., when the concn of urea in the blood equals that in the tissues. In this respect urea is unique as with no other substance investigated is the distribution uniform over blood and tissues. The sudden introduction of urea into the body delays the formation of urea, and may even bring it to a standstill in the first 2 hrs after the injection. The osmotic attraction of the urea injected is greatly reduced by its rapidly leaving the blood. NaCl also leaves the blood so that the osmotic attraction of dissolved substances in the blood is not raised even 3 min after the injection. There is some evidence that except in edematous patients urea injection causes some NaCl to pass out of the body fluids, to be stored in a "dry" form in the salt depots of the body, thus reducing the crystalloid content of the body fluids which had been raised by the injection of urea, but that in most cases of edema, NaCl is liberated from the depots and enters the body fluids. It has been suggested that an important factor in the production of edema is a deficient power of the tissues to hold and bind water, which therefore accumulates in the tissue spaces as edema fluid. The writers suggest that in edema the tissues are also less able to bind NaCl, which also accumulates in the free tissue fluids. H. F. H.

The influence of adrenaline on the immediate variation in alkaline reserve. Role of apnea. Comparative action of formaldehyde and acetylcholine. J. GAUTRELET, D. BENNATI, E. HERZFELD AND L. VALLAGNOSC *Bull soc chim biol* 12, 1100-44 (1930).—Intravenous injection of 0.1 mg adrenaline per kg into anesthetized dogs caused an av. rise in alk. reserve (I) of 15% in 1-3 min., followed by a rapid fall. A condition of apnea is produced but it is not an important factor in relation to I. Apnea produced by intravenous injection of HClO and acetylcholine is not accompanied by a rise in I. Injections of 0.0001-0.01 mg adrenaline produced a distinct rise in I without appreciable apnea. Increased respiration may be due to a sudden release of CO_2 into the blood stream. C. G. KINN.

The action of potassium salts on the Golgi apparatus of the liver cells. M. SHIRASAKA. *Arb Med Univ Okayama* 1, 484-95(1930); *Physiol Abstracts* 15, 588.—The administration of large quantities of KBr and KI to rabbits causes a marked disturbance in the liver cells. KCl, KNO_3 and Na_2SO_4 have only a slight effect. The halogen ions have a destructive action on the Golgi app. of the liver cells and penetrate the bile capillaries, causing an increased bile secretion. It would appear that the Golgi app. bears no relation to the bile capillaries. KNO_3 and Na_2SO_4 have no action on either the Golgi app. or the bile capillaries. G. G.

Action of cholesterol and lecithin on the Golgi apparatus of the nerve cells of rabbits. S. OKADA. *Arb Med Univ Okayama* 1, 503-14(1930); *Physiol Abstracts* 15, 559.—The increase of cholesterol content by lanolin injection favors the development of the nerve cells of rabbits, while the growth of the mitochondria is inhibited. Lecithin has the reverse effect. G. G.

Action of potassium and calcium on the Golgi apparatus and mitochondria of the epithelial cells of the kidney. R. KAMAKURA. *Arb Med Univ Okayama* 1, 515-45 (1930); *Physiol Abstracts* 15, 559.—Repeated injections of $CaCl_2$ bring the Golgi app. of kidney cells into prominence but lessen that of the mitochondria. KCl has the reverse effect. If, however, the injections are carried on for a longer period, the reverse effect of the salts, as noted above, is evident. G. G.

Action of cholesterol and lecithin on kidney tubule cells. R. KAMAKURA. *Arb Med. Univ Okayama* 1, 628-43(1930); *Physiol Abstracts* 15, 560.—Repeated injections of lanolin cause a shrinkage of the cells of the kidney tubules. The Golgi app. is intensified, while the mitochondria sink into the background. Lecithin causes a swelling of the cells, and the mitochondria are intensified. Injections continued over a very long time have the reverse effect in the 2 cases. G. G.

Effect of histamine and peptone on the portal pressure of cats. W. FELDBERG. *Arch exptl. Path. Pharmacol* 140, 156-67(1928); *Physiol Abstracts* 15, 376; *c* *A* 22, 633.—When histamine or peptone is injected into the jugular vein of a cat, the pressure in the portal vein falls secondarily to the fall of general blood pressure. This is sometimes followed by a rise, caused apparently by an increase in the resistance of the liver vessels. When it is injected in the portal vein, there may be a small preliminary rise

in portal pressure which is probably due to the same cause. This constriction of liver vessels is much smaller than that which occurs in dogs, and in some cats it does not occur. G G

The action of pilocarpine and physostigmine on the isolated cat heart after degeneration of the vagus nerve. W. BAANDERDINKA *Arch exp'l Path Pharmacol* 140, 168-73 (1928), *Physiol Abstracts* 15, 361.—Physostigmine and pilocarpine still act on an isolated cat heart (Langendorff prep'n) after degenerative section of the vagi. It is argued that their point of action must be peripheral to the nerve endings. G G

The action of pituitary extract and insulin on the secretion and the blood vessels of isolated suprarenals. M. P. NIKOLAY *Arch exp'l Path Pharmacol* 140, 225-30 (1929); *Physiol Abstracts* 15, 390.—Ringer sol'n was perfused through the suprarenals of cattle and its adrenaline content subsequently det'd on a perfused rabbit ear. It was thus shown that pituitary (posterior lobe) ext. and insulin stimulate the suprarenals to secrete adrenaline. G G

Quantitative investigation on the reversibility of the glucosides of squill as a contribution to the mechanism of the action of heart remedies. W. GAAP *Arch exp'l Path Pharmacol* 140, 315-79 (1929), *Physiol Abstracts* 15, 401.—The action of the glucosides of squills was studied on the heart of *Rana temporaria* (Straub's prep'n). The threshold concns. for systolic arrest are given: scillaren A, $1/1.8 \times 10^6$, scillaren B, $1/3 \times 10^6$, scillardin B, $1/1.5 \times 10^6$. The action is more easily abolished by washing than is the case with digitoxin, and the drug act. quicker and in lower concns. The ease with which the effect could be abolished by washing the heart with various preps. was studied, and the conclusion was reached that the effect depended on a destruction of the drug by the heart rather than on its diffusion out of the muscle. The addn. of dialyzed serum delayed the effect but did not alter the threshold concn. G G

The narcotic action of the vapors of chlorine derivatives of methane, ethane and ethylene. N. V. LAZAREV *Arch exp'l Path Pharmacol* 141, 19-24 (1929), *Physiol Abstracts* 15, 403, cf. C A 24, 2182.—The Cl derivs. of C_2H_6 , C_2H_4 and C_2H_2 were investigated by exposing mice to their vapors in hermetically sealed glass bottles. The min. concn. of the vapor necessary to make the animal (1) lie on its side, (2) lose its reflexes, (3) die, was det'd for each substance. In general it was found that the narcotic action increased with the increase in the no. of Cl atoms. There were some exceptions to this rule, for instance, CCl_4 is less powerful than $CHCl_3$. In these cases it was found that the less powerful comp'd. had a very small soly. coeff. G G

Investigations on the pharmacology and pharmacodynamics of ointments and medications incorporated in ointments. C. MONCORG *Arch exp'l Path Pharmacol* 141, 25-101 (1929), *Physiol Abstracts* 15, 494-S, cf. C A 24, 5863.—I. Ointments and pastes all fall into 2 groups. Either the water is the continuous phase and the fat the disperse phase, or vice versa. The former type is expressed as oil-water in this paper, and the latter as water-oil. The properties of ointments vary according to the type to which they belong, and the type can be varied according to the emulsifying agent used. II. The adsorption and absorption of salicylic acid from ointments was investigated by following its excretion in the urine of men after the various ointments had been applied for 12-24 hours to the skin of the leg. Pasta zinci oxid, vaseline, lanolin, eucerin and physiol. A, B, C, were among the ointments tested. The absorption of salicylic acid was greatest from eucerin (a water-oil ointment) and physiol. C (an oil-water one), but both these have a keratolytic action. The amt. of salicylic acid absorbed is small, and none can be demonstrated until the strength of the ointment reaches 25%. III. S, on the other hand, is absorbed less from physiol. C and eucerin than it is from pasta zinci oxid. During absorption it is reduced, and H_2S can be demonstrated in the blood. IV. After 10% S ointments had been applied for 12 hrs., the total S of the sulfates in the urine was increased. G G

Tests of anthelmintics on the vinegar eel. E. RAY *Arch exp'l Path Pharmacol* 141, 129-41 (1929), *Physiol Abstracts* 15, 407.—A description is given of a series of tests made on the vinegar eel, *Anguilla aceti*, in which its resistance to various substances was compared with that of parasitic worms, earthworms and small fishes. The vinegar eel withstands acids well, while its resistance to alkalis depends on their lipid-sol. Most of the common anthelmintics were tested, and it was found that the vinegar eel is comparable with the nematodes rather than with the cestodes. The anthelmintics tested include morg. substances, alkaloids, org. halogen compds., benzene derivs., *Filix mas* and terpenes. Of the phenol group, it appeared that the higher the mol. wt., the greater the toxicity for vinegar eels, while the toxicity to mammals decreases with increasing mol. wt. Paracymol had the greatest anthelmintic power of the substances

tested. It is suggested that the vinegar eel is a convenient exptl animal for the testing of anthelmintics and gives much better results than the earthworm. G. G.

The curare-like action of tellurites on the frog. R. LANES. *Arch exp'l Path Pharmacol* 141, 142-7 (1929), *Physiol Abstracts* 15, 408.—Small doses of K_2TeO_6 injected into the lymph sac of frogs were found to produce an effect like that due to curare. One to 2 days after the injection the hind limbs were completely paralyzed and the fore limbs partially paralyzed. Stimulation of motor nerves produced no response, but the muscles still responded to direct stimulation. Sensation was intact. The paralysis was a peripheral motor paralysis, but it was not possible to demonstrate the presence of Te in the motor end plates. G. G.

The action of histamine and adrenaline on rabbit ears. E. FLATOW. *Arch exp'l Path Pharmacol* 141, 161-3 (1929), *Physiol Abstracts* 15, 364.—Intravenous injections of 0.025 to 0.05 mg histamine into the vessels of a rabbit ear produced dilatation, while injections of adrenaline, even as small as 10^{-14} mg, always produced a constriction. G. G.

The mechanism of polyphasic action [of drugs]. E. RENTZ. *Arch exp'l Path Pharmacol* 141, 183-227 (1929), *Physiol Abstracts* 15, 400, cf C. A. 23, 908.—A large no. of cases have been collected from the literature in which the most diverse substances produce a biphasic effect on isolated smooth muscle. In some cases both stimulation and inhibition occur successively while the drug is still in contact with the tissue, in others the 2nd phase of the effect is only seen on washing out the drug. The theories which have been propounded to account for the various cases are discussed and found to be inadequate to explain all the cases quoted, which are all thought to be examples of a fundamental property of the tissue. It is pointed out that when biphasic effects are seen in the whole animal, they may be examples of the same phenomenon, and it is unnecessary to postulate special mechanisms such as reflexes to account for them. G. G.

Tolerance of nerves to poisons. A. LÁNCZOS. *Arch exp'l Path Pharmacol* 141, 248-50 (1929), *Physiol Abstracts* 15, 403.—The sciatic nerves of frogs were immersed in solutions of urethan, and the strength of current required to stimulate them was measured at intervals of 15 min. It was found that a nerve left in a soln of urethan became narcotized and subsequently recovered while still in the soln. Also a nerve which had been previously immersed in a weak soln of urethan was not narcotized by a soln which would completely narcotize an untreated nerve. Motor-nerve endings, on the other hand, could not be shown to acquire any tolerance to urethan. G. G.

Pharmacology of the rare-earth metals. II. Yttrium. H. STEDLE AND M. DING. *Arch exp'l Path Pharmacol* 141, 273-9 (1929), *Physiol Abstracts* 15, 408, cf C. A. 24, 4839.—Yt ppts albumin, inhibits the action of yeast cells, paralyzes frogs and causes at 1st excitement and dyspnea, and later paralysis and death in mice. It stops an isolated frog heart in diastole. In general Yt may be said to resemble Al and Ce in its action, and its toxicity is about the same as that of Ce, except in its local action, in which it is more powerful than Ce. G. G.

Sensitivity of the rabbit heart to strophanthin after feeding with irradiated ergosterol. R. MAYCKE. *Arch exp'l Path Pharmacol* 141, 280-91 (1929), *Physiol Abstracts* 15, 401.—With normal rabbits, a concn of strophanthin of 10^{-4} in the perfusion fluid invariably made the heart stop in systole. Hearts of rabbits which had been fed with irradiated ergosterol, but in which no pathol changes could be demonstrated, behaved like normal hearts. Hearts of rabbits with pathol changes took a longer time to stop in systole, and before that stopped periodically in diastole. They also showed no dilatation of the coronary arteries during the perfusion, whereas in normal hearts a slight dilatation occurred. G. G.

Disposal of digitalis and its action in warm-blooded animals. II. Extra-cardiac consumption of digitalis and the conditions of glucose combination in the heart. H. WEESE. *Arch exp'l Path Pharmacol* 141, 329-50 (1929), *Physiol Abstracts* 15, 400, cf C. A. 24, 4840.—It has been shown that only a small proportion of the active principles of digitalis is stored in the heart on intravenous injection. An attempt is here made to trace the fate of the rest. Various organs were perfused by a cat's heart-lung prep'n, and the digitalis (or strophanthus) injected into the arterial cannula going to the organ. The lethal dose for the heart was known. The amt. injected in excess of this before the heart stopped was presumed to be stored in the tissue. It was thus shown that about half the min. lethal dose could be taken up by the voluntary muscles, but the heart and kidneys absorbed the largest quantity in proportion to their wt. By working with almost poisoned hearts it was shown that the different organs allowed no digitalis to pass them till they were nearly sat'd. Uptake of digitalis by the isolated

serum. J LA BARRE AND J LOUIS WODON *Compt rend soc biol* 104, 111-2(1930); *Physiol Abstracts* 15, 414—Gunea-pig uterus *in vitro* was treated with sensitizing serum in Locke's fluid contg tissues from animals sensitized some weeks previously. The hypertonic contractions indicated anaphylactic shock. This was produced in 1 cornu, the other serving as control. The addn of Mg salts effected immediate cessation of the symptoms of shock. The salts produce tissue desensitization or render the uterus insensible to stimulation. The hyposulfite is the most suitable salt. G. G.

Action of hypertonic sodium chloride on intestinal motor activity. L BOUSSIER AND P FABRE *Compt rend. soc biol* 104, 462-5(1930), *Physiol Abstracts* 15, 452—The record of intestinal contractions taken from a dog shows increase of activity when 10-20 cc of a 30% soln of NaCl is injected into a vein. This stimulating influence of hypertonic NaCl is const. G. G.

Action of vagotonin on glucemia. T BRIEL G FLCHS, D SANTENOISE AND M VIDA COVITCH *Compt rend soc biol* 104, 763-70(1930), *Physiol Abstracts* 15, 594—The vagotonic substance secreted by the pancreas and isolated from insulin in pancreatic exts by frequent washing with alc. possesses some hypoglucemic action. Its action on the blood sugar is effected through the vagus nerve, thus it differs from insulin, the action of which is direct. G. G.

Anticoagulating action of sulfarsphenamine. E ZUNE AND M A CAMACHO *Compt rend soc biol* 104, 703-4(1930), *Physiol Abstracts* 15, 571—The clotting of recalcified oxalated plasma from the rabbit was found to be retarded by the action *in vivo* of sulfarsphenamine. The effect is essentially due to a combination or complex of the arsenical compd with fibrinogen. The arsenical compd seems also capable of combining with and hindering the action of thrombin, and it prevents the conversion of proserozyme to serozyme. G. G.

Action of hypertonic sodium chloride in intestinal occlusion. L BOUISSET AND P FABRE *Compt rend soc biol* 104, 847-50(1930), *Physiol Abstracts* 15, 452—By injection of hypertonic NaCl at varying intervals after occlusion of the small intestine in the dog it has been found that the muscular tone passes through 2 distinct phases. In the 1st the wall is sluggish but ready to react violently under the influence of the excitant. The pressure within the lumen rose from 2.5 cm Hg to 8 cm after the injection. In the 2nd phase atony is pronounced, and the reactions are weak or neg. It requires 5 or 6 days of occlusion to show the complete change. G. G.

Action of camphor and its derivatives on the isolated intestine. H BUSQUET *Compt rend soc biol* 104, 869-72(1930), *Physiol Abstracts* 15, 550—Camphor and its sol. derivs, campho-sulfonate and campho-carbonylate of soda, act as excitants of the central nervous system and as cardiac tones. Pure camphor exercises a marked inhibitory action on the isolated intestine, but the derivs mentioned cause a distinct increase of amplitude of the contractions. The sedative action of camphor must have been suspected when it was introduced into paregoric, but it is essential, in prescribing it for intestinal spasm, to bear in mind the difference of action from that of the derivs. G. G.

Action of vasomotor substances on albuminuria. A HANNS *Compt rend soc biol* 104, 877-8(1930); *Physiol Abstracts* 15, 529—Injection of adrenaline in one subject of Bright's disease was invariably followed by diminution of albumin in the urine which lasted some hrs. The excretion of water in excess may have accounted for the change. In a 2nd case the drug had the opposite effect on the albumin and none on the excretion of water. Postural influences on the amt of urine and the albumin content were different in the 2 cases, and it is uncertain if the differences in the action of the drugs were related to this influence. G. G.

Absorption capacity of dental pulp. A E ORIGLIO *Rev soc Argentina biol* 6, 79-98(1930); *Physiol Abstracts* 15, 551—The 1st molars of dogs were trephined, and substances were introduced into the cavities. KI (tested in blood, urine and saliva) and phenolsulfonephthalein (in 24 hrs 4-26% passes into the urine) are easily absorbed. In doses varying between 1 and 20 mg the following substances are absorbed: apomorphine (vomiting), strychnine (convulsions), atropine (mydriasis and vagus block), cocaine (death), nicotine (death), cyanides (slight symptoms). Absorption of the following substances could not be proved: venoms of *Lachesis alternatus* and *Naja tripudians*, insulin and adrenaline. Perforation of the tooth produces a slight hyperglucemia. G. G.

Chemical structure of insulin. Study of some synthetic sulfur compounds having hypoglucemic action. C RUIZ, L SILVA AND L LIBENSON *Rev soc Argentina biol* 6, 134-41(1930); *Physiol Abstracts* 15, 530—Since S is found in insulin, the hypoglucemic action of S compds was tested in rabbits *per os* and by injection. Cyclodecamethylene-

thiourea is insol and inactive Thiophenol, thioacetic acid and 4- or 5-methyl-2-thioimidazole have some activity (blood sugar falls from between 0.099 and 0.132 to between 0.072 and 0.08%) G G

Calcium and magnesium salts and the healing of wounds. G MAZZACUVA Riv patol sper 5, 327-42(1930), Physiol Abstracts 15, 483—The formation of scar tissue on superficial wounds is hastened by the subcutaneous injection of small quantities of Ca and Mg chlorides G G

Suppression of glucose combustion by adrenaline administration. A. R. COLWELL Endocrinology 15, 25-33(1931)—Cats were anesthetized by means of amytal, and the effects of continued administration of 2 g glucose per hr and 0.001 mg of adrenaline per min per kg upon the total and non protein respiratory quotients, and blood sugar, were detd After 3-5 hrs of glucose and adrenaline, the non protein quotient reached a level near 0.7 and did not change appreciably until the adrenaline was discontinued, whereupon it rose to normal in 4 hrs Therefore, adrenaline interferes with, and may even abolish, glucose combustion W D LANGLEY

The toxicity of amyl nitrites. OTTO GERHARDT Chem-Ztg 55, 128(1931)—G describes the toxic effects exerted upon himself and his associates by amyl nitrite in the process of its prepn experimentally and on a large scale The toxic products become manifest in the early stages of the nitration of the amyl alcs and cause headache, increased blood pressure, general dizziness, noises in the ears, etc The effects are cumulative, hence the danger of repeated slight intoxications B S LEVINE

Chlorophyll and its nephrohepatic extracts in experimental anemia. CARLOS A SAGASTUME and JOSÉ A PEZZANI Rev facultad cien quim (Univ La Plata) 7, Pt II 7-22(1930)—Chlorophyll possesses interesting pharmacodynamic properties, the study of which may lead to important therapeutic applications In exptl anemias studied S and P found chlorophyll acting as an agent regenerating erythrocytes and hemoglobin B S LEVINE

The combined action of ultra-violet rays and neosarsphenamine on trypanosoma equiperdum. G GIEMSA AND V ELLENBOGEN Z Immunitäts 67, 78-85(1930)—Neosalsarvan in amts slightly less than the curative dose injected into mice infected with trypanosoma equiperdum causes all the organisms to disappear from the blood in 2-10 days only to reappear within several days Radiation of the mice after the injection of the salsarvan does not alter this course of events, contrary to the findings of Roskin and Romanowa (C A 24, 5667) JULIAN H LEWIS

The combined action of ultra-violet rays and a series of trypanocidal substances on trypanosomes. GA ROSKIN, A BICHOSKAYA AND S SHEVCHIKOVA Z. Immunitäts 67, 91-4(1930) cf following abstr—Among several trypanocidal substances novarsolan was the only one whose action was enhanced by exposure of the infected mice to ultra violet rays after injection of the drugs JULIAN H LEWIS

The method of combined therapy with arsenicals and ultra-violet light. GA ROSKIN Z Immunitäts 69, 240-3(1930), cf 2 preceding abstrs—G and E failed to confirm the observation of Roskin that mice infected with trypanosomes and injected with ineffective doses of novarsolan are sterilized if the mice are exposed to ultra violet light after injection of the novarsolan This failure is attributed to the fact that Giemsa and Ellenbogen failed to remove the hair of the mice before irradiation J H L

The analysis of the combined action of arsphenamine and ultra-violet light in protozoan infections. G ROSKIN AND K ROMANOVA Z Immunitäts 67, 94-101(1930), cf 2 preceding abstrs—Serum of mice exposed to ultra violet light can change a non curative dose of novarsolan into a curative one Serum of radiated mice alone has no trypanocidal action JULIAN H LEWIS

Analysis of the combined therapy with drugs and ultra-violet light. GR ROSKIN Z Immunitäts 69, 473-83(1931)—In the serum of mice irradiated with ultra violet light there is a factor transferable to other mice that enhances the effect of neosarsphenamine in the treatment of trypanosome infections in mice This factor is destroyed by heating the serum at 56° for 1½ hr and by irradiating the serum in vitro with ultra violet light It is adsorbed by kaolin The factor arises from the mesenchyme of the skin and the entire reticulo endothelial system JULIAN H LEWIS

Further chemotherapeutic studies of a new benzene arsenic acid derivative (arsenic preparation 4002). GIEMSA Z Immunitäts 69, 86-99(1930)—Theoretically the arsenic acid compds should be better adapted to the treatment of trypanosomal diseases than the arseno compds, since they are more easily diffusible and do not combine with the colloids of the blood This assumption is borne out by expts which show that As prepn 4002, a benzene arsenic acid deriv, cures trypanosome infections in mice better

than do atoxyl and tryparsamide Syphilis in rabbits is also cured by this prepn given either parenterally or orally JULIAN H LEWIS

Can biological tests of toxicity of arsenobenzol derivatives be replaced with chemical tests? H BAUFER AND M ROTHERMUNDT *Z. Immunitäts* 69, 213-9(1930)—The chem method of Kielbasinski (*C A* 24, 5868) does indicate the full toxicity developed in arphenamine through oxidative processes JULIAN H LEWIS

Changing arphenamine fastness. HEINRICH CITRON *Z. Immunitäts* 69, 464-71 (1931)—Repeated injections of $\text{Na}_2\text{S}_2\text{O}_4$ into mice infected with arphenamine fast trypanosomes permanently destroyed the arphenamine fastness The low toxicity of $\text{Na}_2\text{S}_2\text{O}_4$ permits the use of this principle in patients JULIAN H LEWIS

The behavior of lead in the animal organism II. Tetraethyl lead. ROBERT A KILBOE AND FREDERICK THAMANN *Am J Hyg* 13, 478-98(1931), cf *C A* 22, 2800—The distribution of tetraethyl lead in rabbits following absorption by the skin of the abdomen was detd over short and long periods Any excess was washed off after 60-90 min Inhalation of the lead was prevented Tissues, bones and blood were analyzed by steam distn of the tetraethyl lead from them From pure tetraethyl lead the absorption rate is rapid From 0.1% in gasoline (the percentage in treated gasoline) absorption was negligible through the skin During rapid absorption the distribution in tissue corresponds to that of an oil sol material The compd is rapidly decompd by tissue, so that after a period of 1-2 weeks the lead deposit resembles that of a water sol lead compd The nervous system is very susceptible to lead absorption excretion follows quantitatively that of water sol compds G H W LUCAS

Brain-water movements during anesthesia II G BARBOUR *Science* 73, 346-7 (1931)—As a possible explanation of anesthesia, the dehydration of nerve cells has been suggested, but expl proof of this is a difficult undertaking Anesthetic substances vary in their effect on the water content of the brain as a whole and of certain portions of it In some ether and morphine expts a hydration was evident and again withdrawal of morphine from young rats and dogs resulted in hydration Other expts have indicated in the first 2 hrs after morphine the cerebrum is dehydrated but the medulla is hydrated This increase in ratio of medulla H_2O to cerebrum H_2O has been found with amytal and others Ether excitement prolonged for half an hour did not cause an increase in the ratio The ratio for normal rats is 0.022 ± 0.003 morphinized rats showed 0.039 ± 0.007 Possibly the colloidal conditions of the cells are altered in such a way as to extrude water G H W LUCAS

Effect of continuous intravenous injection of adrenaline on the carbohydrate metabolism, basal metabolism and vascular system of normal men. C F COAR AND K. W. BUCHWALD *Am J. Physiol* 95, 71-8(1930)—Adrenaline was injected intravenously into normal men for a period of 30 min, 0.00003 mg per kg per min was the lowest rate which still caused a rise in blood pressure, blood lactic acid and basal metabolism. 0.000025 mg per kg per min caused an increase in pulse rate, respiration and blood sugar J F LYMAN

Effect of adrenaline on sugar utilization in animals under amytal anesthesia. G T CORI *Am J Physiol* 95, 285-91(1930)—Amytal anesthesia, which in itself has but a slight effect, greatly intensifies the depressive action of adrenaline on glucose utilization Glucose utilization by rats was as follows unanesthetized 98%; unanesthetized + adrenaline 89%, amytal anesthesia controls 91%, amytal + adrenaline given subcutaneously 40%, amytal + adrenaline given intravenously 27% J F LYMAN

The effect of various anterior pituitary preparations upon basal metabolism in partially thyroidectomized and in completely thyroidectomized guinea pigs. W J STUBBS AND R. S SMITH *Am J. Physiol* 95, 396-402(1930)—Armour pituitary tablets caused a marked rise in the basal metabolism of guinea pigs, independently of the thyroid glands An acid ext of the anterior lobe of the pituitary caused a rise in metabolism only when the thyroids were intact and seemed to depend on stimulation of the thyroid by the pituitary ext. J F LYMAN

Depletion of muscle sugar by adrenaline. II. F BISCHOFF AND M LOUISA LONG *Am J. Physiol* 95, 403-11(1930), cf *C A* 24, 3561—Adrenaline injected subcutaneously or by the continuous intravenous method (at least 4 times the min hyperglucemic dose) caused a fall in muscle sugar J F LYMAN

β -Tetrahydronaphthylamine hyperthermia and fat metabolism. JEAN J. BOUCK AERT AND E SOLOMON *Am J. Physiol* 95, 417-21(1930), cf *C A* 24, 5065—The increased metabolism observed after injecting β tetrahydronaphthylamine is independent of the exchanges of fat between the liver and the other tissues It is probable

that the excess material metabolized after taking the drug is protein after the carbohydrates have been depleted. J. F. LYMAN

The action of histamine on the motility of different parts of the intestinal tract. MARGARET I. MACKAY *Am J Physiol* 95, 527-530(1930)—Intravenous doses of 0.25 to 0.50 mg. histamine activate a quite characteristic motor response in different parts of the intestine, most marked in the ileum and decreasing toward the duodenum. Local application of histamine to the outside of the intestine provoked the typical reactions, but injection of the drug into the lumen of the bowel did not activate a motor response. J. F. LYMAN

The effect of a solution of acacia in restoring diminished body fluid. M. A. WALKER AND N. M. KEITH *Am J Physiol* 95, 561-72(1930)—Conclusion: Acacia has the property of holding the injected fluid in the circulation for a considerable time. It does not inhibit renal excretion of water. J. F. LYMAN

The heart rate after sympathectomy and vagotomy, and the blood sugar as affected by posterior hypophyseal extracts (pitressin and pitocin). Z. M. BACQ AND S. DWORKIN *Am J Physiol* 95, 605-13(1930)—In normal cats pitressin, intravenously injected, causes a prompt reduction in heart rate. Pitocin has little effect. In vagotomized cats, pitressin reduces the heart rate, but to a level far above the basal rate for a normal animal. After removal of both sympathetic chains, pitressin slows the heart when the vagi are intact, but causes marked acceleration when these nerves have been cut. In normal animals both pitressin and pitocin, given intravenously, cause a prompt rise in blood sugar which may persist as long as 2 hrs. In sympathectomized cats the glucemic effect of pitressin is somewhat less than in normal animals, but not markedly so. J. F. L.

The action of parathyroid extract in sympathectomized animals. Z. M. BACQ AND S. DWORKIN *Am J Physiol* 95, 614-9(1930)—The absence of the central sympathetic nervous system does not appreciably influence the response of cats and dogs to parathyroid ext. J. F. LYMAN

The effect of adrenaline upon the nitrogen metabolism of rabbits. O. WATKINS AND G. VAN S. SMITH *Am J Physiol* 96, 28-34(1931)—Intravenous, subcutaneous or intraperitoneal injections of 0.1 mg. of adrenaline cause marked rises in blood sugar and urea of the blood in rabbits. Since it was shown that the excretion of urea by the kidney is not interfered with, the rise in blood urea is a result of some catabolic action of adrenaline upon protein metabolism. The intravenous injection of glucose after adrenaline injection does not affect the rise in blood urea, while glucose and insulin together exaggerate and prolong the urea rise. J. F. LYMAN

The temporary control of post-operative tetany in thyro-parathyroidectomized dogs by the administration of thyroid hormone. M. M. KUNDE, R. M. O'LEND AND R. KERN *Am J Physiol* 96, 45-53(1931)—In approx. 35% of 31 dogs, tetany, following the removal of the thyroid and parathyroid glands, was temporarily allayed by administering dried thyroids or thyroxine. In unoperated dogs, moderately large doses of thyroid or thyroxine caused an elevation in acid sol. P of the blood to as much as 100% above normal. This high blood P persisted after removal of the thyroids and parathyroids regardless of the decrease in serum Ca. J. F. LYMAN

The action of atropine and adrenaline on gastric tonus and hypermotility induced by insulin hypoglycemia. R. L. WILDER AND F. W. SCHULTZ *Am J Physiol* 96, 54-8(1931)—The increased tonus and hypermotility of the stomach, which takes place when the blood sugar is markedly lowered following insulin is due, in part at least, to increased irritability of the vagus motor nerves to the stomach. Atropine and adrenaline lower this increased tonus without raising the blood sugar. J. F. L.

Thiocyanate contracture in skeletal muscle. Permeability of muscle. E. GELHORN *Am J Physiol* 96, 203-13(1931)—The immersion of a frog sartorius muscle in a Ringer soln. contg. NaSCN in a concn. which is below that causing contracture has no direct effect upon the muscle but gives rise to a contracture when this soln. is replaced by common Ringer soln. The effects of various salts on this indirect contracture were studied. Conclusion: The velocity with which NaSCN leaves the muscle cell detrs. the strength of indirect contracture. J. F. LYMAN

The physiological action of glyoxals. A. HYND *Proc Physiol Soc., J Physiol* 70, 12-11(1930)—Solid hydroxymethylglyoxal, prepd. by the method of Evans and Waring (*C. A.* 20, 3692) is as a rule, contaminated with an unstable, highly toxic by-product, possibly a thiol deriv. There is no essential difference in the effects produced by the monomeric and dimeric forms of hydroxymethylglyoxal. There is a close analogy between the effect of glyoxal and of that of hydroxymethylglyoxal, but no similarity between that of the latter and glucosone. J. F. LYMAN

"Dial" as an anesthetic for surgical operations on the nervous system. J. F. FULTON, E. G. T. LODELL AND D. McK. RITCH. *Proc. Physiol. Soc., J. Physiol.* 70, 1113 (1920), cf. *C. A.* 25, 2292.—"Dial Ciba" is a liquid prepn. of diallylbarbiturate acid (10%) combined with small amts. of urethan and monoethylurea (to assist soln.) With this anesthetic the spinal reflexes remain active even in profound anesthesia. In exptl. surgery of the nervous system the drug has 2 important uses: (1) It appears to constrict small blood vessels of the brain so that removal of the cerebellum or cerebral hemispheres can be carried out almost bloodlessly without occlusion of any of the great arteries to the head. (2) Animals remain quiescent for a period of 12-36 hrs. and thus give opportunity for satisfactory healing of the incision. J. F. LYMAN

The influence of adenosine and related compounds upon the coronary arteries in the perfused rabbit's heart. A. N. DRURY AND A. M. WEED. *Proc. Physiol. Soc., J. Physiol.* 70, 1118 (1920), cf. *C. A.* 24, 2786.—Adenosine dilates the coronary arteries of the heart. Yeast adenylc acid and muscle adenylc acid have about 66% of the effect of adenosine and inosine acid; adenine, guanosine and NaNO_2 are about 33% as effective as adenosine, while thymus nucleic acid and yeast nucleic acid have an inconstant effect; a slight constriction is sometimes produced. J. F. LYMAN

The selective vasoconstrictor action of pituitary pressor extract. G. A. CLARK. *J. Physiol.* 70, 53-9 (1920).—Pituitary pressor ext. (pitressin) intravenously injected in cats caused a redistribution of blood so that the intestine contained much less and the muscles more; the skin and connective tissues usually contained less, but if the blood pressure was greatly increased, might contain more. Pitressin exerts its most powerful vasoconstrictor action on intestinal vessels and its weakest action on the vessels of voluntary muscles, while those of skin and connective tissue occupy an intermediate position. J. F. LYMAN

The effect of adrenaline on muscle glycogen. A. B. CORKHILL AND H. P. MARKS. *J. Physiol.* 70, 67-83 (1920).—Adrenaline alone and adrenaline and insulin together were administered to decapitated cats whose viscera had been removed, and the subsequent changes in carbohydrate content of blood and muscles observed. Conclusion: Adrenaline causes discharge of glycogen from resting muscle. Part of this glycogen appears as lactic acid and another part as some unknown substance which is non-reducing or has a lower reducing power than glucose does, or it fails to be extd. for analysis by the methods used. A similar discrepancy was noted when insulin and adrenaline acted together. J. F. LYMAN

The production of hypoglycemia in rabbits by oral administration of precipitated insulin. H. N. MCKINSTRY. *J. Physiol.* 70, 182-3 (1930).—The phosphotungstic acid ppt. of crude insulin, administered to fasting rabbits by mouth, lowered blood sugar to the convulsion level within 1 to 2 hrs. Beneficial effects have been noted with diabetic patients after its administration. It appears to have no toxic effects. J. F. LYMAN

The effect of insulin and other endocrine extracts on the cholesterol content of tissues. E. N. CHAMBERLAIN. *J. Physiol.* 70, 441-8 (1930), cf. *C. A.* 24, 4822.—Overdosage of rabbits with insulin caused a marked decrease in the cholesterol content of the suprarenals, but this effect was not so pronounced if the hypoglycemia was prevented by simultaneous administration of glucose. Neither insulin, thyroid nor pituitrin produced any material changes in the cholesterol content of the blood, liver or spleen. It is suggested that the synthesis of cholesterol by the adrenals is controlled by the internal secretion of the pancreas. J. F. LYMAN

Intraocular pressure. I. The action of drugs on vascular and muscular factors controlling the intraocular pressure. J. COLLE, P. M. DUKE ELDER AND W. S. DUKE-ELDER. *J. Physiol.* 71, 1-30 (1931).—The effects of adrenaline, pituitrin, histamine, choline, nicotine, atropine, physostigmine and curare on intraocular pressure were detd. All the changes observed are compatible with the theory that the aq. humor is a dialyzate. J. F. LYMAN

The influence of creatine on the excretion of phosphates by the kidney. MARION BROWN AND C. G. IMBRIE. *J. Physiol.* 71, 222-8 (1931).—The administration of creatine to cats is followed by a temporary fall in the output of phosphates by the kidney, which is most marked when the larger amts. of creatine are retained. The concn. of P in the blood may be above the normal when phosphates are absent from the urine. J. F. LYMAN

Experimental investigation on the fate of bilirubin introduced into the blood vessels. II. The excretion of bilirubin from the liver. I. The localization of the excretion of bilirubin of the liver. S. SAIKI. *Japan J. Gastroenterology* 3, 1-13 (1931), cf. *C. A.* 25, 1277.—When bilirubin is injected intravenously into rabbits, it is excreted into the

ble by the parenchymatous cells of the liver. This was shown by blocking the reticulo-endothelial cells of the liver with colloidal silver and India ink injections. Since neither retarded the excretion of bilirubin, these cells do not function in such an excretion. Injury to the parenchymatous cells by injecting the rabbits with hepatotoxins decreased the rate of excretion of bilirubin from the blood, so these cells must be responsible. C. M. McCAY

Experimental study on the influence of poisonous gases upon the pigment-excreting function of the liver and the kidneys. II. Cyanogen. H. YAGATA. *Japan J. Gastroenterology* 3, 80-4(1931) and C. A. 25, 351.—In rabbits the pigment-excreting function of the liver and kidneys is injured by HCN poisoning. The kidneys suffer greater injuries, which increase the dye-excreting function of the liver. Similar expts were carried out with rabbits after the administration of such liver poisons as hepatotoxin, CHCl_3 , and P or kidney poisons such as cantharidin, K_2CrO_4 , or U nitrate. The common liver poisons seem to injure the kidneys so severely when acting with HCN that the liver dye excretion increases. Kidney poisons alone increase the dye excretion of the liver but accompanied by HCN the liver excretion is decreased. C. M. McCAY

The effect of pilocarpine in a case of sympathicotonia in childhood. FLORNE M. LANDIS AND J. C. GRIFFING. *Am. J. Diseases Children* 39, 1022-30(1930).—Hyperactivity of the sympathetic portion of the autonomic nervous system, termed sympathicotonia, appears to be reduced following the intramuscular administration of physostigmine or the intravenous administration of pilocarpine. The symptoms of the disease appear to be aggravated by atropine or thyroid ext. F. R. MAIN

The toxicity of α -butyloxy-cinchonic acid diethylethylenediamide hydrochloride (nupercaine). W. R. BOND AND N. BLOOM. *J. Lab. Clin. Med.* 16, 447-51(1931).—The min. fatal dose of nupercaine, injected subcutaneously, is approx. 20 mg. per kg. for guinea pigs and 25 mg. per kg. for dogs. On intravenous injection, it is 2.5-3 mg. per kg. for dogs. It is rapidly absorbed from the nasal and buccal passages, but very slowly from the bladder or vaginal canal and appears to be rapidly detoxicated *in vivo*. F. R. MAIN

Purine elimination in the rat. PIETRO RONDONI. *Arch. sci. biol. (Italy)* 15, 579-85(1930).—Studies were made on the elimination of uric acid and allantoin in the urine of the rat following the administration of purine free diets and of hyperpurine diets (horse liver and calf thymus). The rat has a high protein metabolism which is further increased following the prolonged administration of a thymus diet. Uric acid is essentially the end product of the endogenous purine metabolism, while following the administration of hyperpurine diets allantoin in marked amts. is also eliminated. The transformation of uric acid into allantoin is a sort of protective action against the accumulation of uric acid and is brought into play by the organism which has been subjected to a forced purine metabolism. Repeated applications of CHCl_3 seem to reduce the tendency of the rat to transform excess uric acid. PETER MASCECI

Hematoporphyrinuria and intoxication by fenchone plus sulfonal and camphor oxime plus sulfonal. A. BONAYNE. *Boll. soc. ital. biol. sper.* 5, 1106-7(1930).—Functional alterations in the liver were induced by administering to rabbits under const. diet, sulfonal (I), sulfonal and fenchone (II), sulfonal and camphor oxime (III). The urine of all the animals had a normal color, no albumin was present. In regard to hematoporphyrin, the animals receiving (I) gave results similar to those previously reported by B., while no hematoporphyrin was found in the urine of animals receiving (II) or (III). PETER MASCECI

The influence of certain drugs on pulmonary ventilation in normal and phrenectomized dogs. A. JAFFELL. *Boll. soc. ital. biol. sper.* 5, 1126-8(1930).—Pulmonary ventilation was detd. in (a) normal and (b) phrenectomized dogs after the administration of certain drugs. Morphine HCl in doses of 0.003 g. per kg. had a depressive action in the normals both as to respiratory frequency and pulmonary ventilation; in the phrenectomized animals the action was more accentuated and more persistent. Heroin-HCl in doses of 0.002 g. per kg. produced in (a) a slight diminution of frequency and a marked diminution of pulmonary ventilation, in (b) the injection was followed by an increase in frequency and in ventilation. Eucodale, in doses of 0.002 g. per kg. produced in both (a) and (b) a marked and prolonged diminution of frequency and ventilation. The other drugs studied were: atropine sulfate, belladoline, caffeine, lobeline HCl, adrenaline HCl, synthetic ephedrine and ergotamine tartrate. The general conclusion is drawn that the respiratory app. of the bilateral phrenectomized dog is relatively sufficient under ordinary conditions but does not function sufficiently when, through the agency of drugs, one seeks to raise its activity. PETER MASCECI

The influence of physostigmine on the contractile activity of the spleen. P. TES-

TONI *Boll soc ital biol sper* 5, 1128(1930) —Physostigmine produces a polyglobulia which is essentially the concomitant of the degree and duration of the splenic contraction

PETER MASUCCI

Thallous acetate and hematoporphyrin. P TESTONI *Boll soc ital biol sper* 5, 1130(1930) —While in P's poisoning hematoporphyrin appears in the urine, in acute or chronic TIOAc poisoning, hematoporphyrin is absent, although the urine is highly pigmented

PETER MASUCCI

Studies on the calcium-fixing power of the lung subjected to artificial pneumothorax.

I. The fixation of calcium in the collapsed lung. RENATO PACHOLLI *Boll soc ital biol sper* 5, 1184-6(1930) —Unilateral pneumothorax treatment (in rabbits) does not cause a greater fixation of Ca in the collapsed lung in contrast to the non-collapsed lung

II. The action of a protracted calcium treatment on the fixation of the metal in the collapsed lung. *Ibid* 1186-9 —Rabbits were subjected to unilateral artificial pneumothorax and one week later injected daily with 10 cc Ca gluconate soln (10%) for a period of 12-17 days. The amt of Ca was then detd in the collapsed and normal lung. There was an increase of Ca in both lungs compared with the value found in the lungs of untreated normal rabbits. The increase was 4.08% mg for the non-collapsed and 10.03% mg for the collapsed lung

PETER MASUCCI

Cerebral lipoids in relation to hypno-anesthesia. IVO NOVI *Boll soc ital biol sper* 5, 1190-2(1930) —Crit observations

PETER MASUCCI

Studies on phlorhizin. III. The action of phlorhizin on the isolated heart of the tortoise. P S ISRAEL *Boll soc ital biol sper* 6, 8-10(1931) —The heart was perfused with Ringer soln contg 0.1% phlorhizin, 15-20 min afterward the ventricular contractions became less and diminished in amplitude, after $1\frac{1}{2}$ hr they stopped completely. With concns less than 0.1%, the frequency of the contractions diminished without appreciable changes in amplitude. The contractions stopped after 5-6 hrs of perfusion. Phlorhizin affected the auncles to a lesser degree. After washing the heart repeatedly with Ringer soln alone, the contractions became normal 30 min later. The heart was then perfused with Ringer soln contg 0.01% atropine sulfate. The introduction of phlorhizin induced a phenomenon similar to that already described. When the atropine and phlorhizin were introduced simultaneously in the perfusion liquid, there was a marked decrease in the frequency of contractions. Adrenaline 0.05% and phlorhizin, when added simultaneously to the perfusion liquid, caused no changes in the contraction of the ventricle or auncles. IV. The influence of phlorhizin on imbibition. *Ibid* 10-2 —The gastrocnemius and liver of the frog were immersed in Ringer soln contg 0.10-0.15% phlorhizin and weighed at intervals of 1, 2, 4, 6 and 10 hrs. Phlorhizin tended to increase the imbibition after the first 3 or 4 hrs, the increase was 6-10% greater than that of the controls immersed in Ringer soln alone. The imbibition was more marked with the liver. The increase in wt was about 20% greater than the controls

PETER MASUCCI

The action of synthetic thyroxine on the electric excitability of the vagus nerves. V. G BARONE *Boll soc ital biol sper* 6, 18-20(1931) —The threshold of excitability of the vagus nerves of the dog was detd before and after the injection of the drug. The intravenous injections of 5-6 mg thyroxine were well tolerated by the animals and caused neither a change of arterial pressure nor variations in the frequency of the pulse or respiration. Even with repeated injections, there was always an increase of the elec excitability of the vagus nerves lasting 30-45 min. The increase in excitability was more marked with doses of 1-2 mg, it was distinct but not so marked with doses of 5-6 mg. These results show that in regard to the vagal excitability, this analogy between the pharmacol action of thyroxine and thyroid substances confirms the point that the former has the property of a thyroid hormone

PETER MASUCCI

The effects of intravenous injections of glycine in normal and diabetic individuals. A COSTA AND V G BARONE *Boll soc ital biol sper* 6, 21-4(1931) —The hyperaminoacidemia obtained by intravenous injections of 5 g glycine is not accompanied by any appreciable modifications of the glucose titer either in the diabetics or in the normals. The hyperaminoacidemia observed after the ingestion of proteins cannot be held as the cause of the hyperglucemia in diabetics. The amino acid curve is not noticeably different in the normals than in the diabetics. The clinical observation that certain proteins have a marked hyperglucemic power in diabetics in contrast to other proteins, indicates possibly that the hyperglucemia induced by the ingestion of proteins and occurring simultaneously with the hyperaminoacidemia is caused by the presence of special amino acids, among which glycine must be excluded

PETER MASUCCI

Variations in toxicity obtained by certain respiratory drugs after cutting the vagus nerves. EMILIO TRABUCCII *Boll. soc. ital. biol. sper.* 6, 49-52(1931). —The purpose

was to det. whether after cutting the vagus nerves a difference might be detected in the immediate m l d of certain drugs which kill by arresting the respiratory function. The expts. were made on rabbits using MgCl , SrCl , morphine-HCl and NaOH. The solns. were injected into the jugular vein. A marked difference in the m l d was revealed between the animals having the cut and uncut vagi. The percentage diminutions of the m l d in animals having the cut vagi were: MgCl , 38, SrCl , 37, NaOH 40 and morphine HCl 33. The conclusion is that after cutting the vagi, the respiratory centers are much more sensitive to the paralyzing action of certain drugs, and probably in general, to all drugs which kill by depressing the respiratory centers. P. M.

Apnea from atropium. I MILIO TARBUCCII *Boll. soc. ital. biol. sper.* 6, 53-6 (1931).—The injection of SrCl solns. of various strengths into rabbits produces typical periods of apnea. The bilateral cutting of the vagus nerves does not impede the apnea. The apnea is probably of central origin, since Sr has little action on the neuro-motor terminal junctions. PETRA MASUCCI.

Calcium therapy. Calcium pyruvate. ANGLO PUGLIESE AND DOMENICO ROMOLI-VENTURI *Vinetta med.* 1931, 1, 366-71.—Ca pyruvate is easily sol in water, its aq. solns. are stable, since neither boiling nor long exposures to ice box temps. produces a ppt. The salt is easily absorbed when injected intramuscularly or subcutaneously, and on reaching the circulation Ca ions are rapidly liberated. The substance is non-toxic and may be administered intravenously in high doses to man, it favorably influences the cardiac function without any noticeable changes in the pressure or in respiration, stimulates the automatic contractions, and increases the muscular tone of the intestine and in general of smooth muscle tissue. PETRA MASUCCI.

The effect of adrenaline on the circulating blood volume in individuals with normal and enlarged spleens and after splenectomy. CHIH-SHUN YANG AND HSIAO-CH'EN CHANG *Chinese J. Physiol.* 4, 21-9 (1930), cf *C. A.* 22, 2622.—The injection of adrenaline into human subjects with normal or enlarged spleens, following the inhalation of CO , resulted in an increase in red cells and CO content of the blood but no change in the computed blood vol. Splenectomized subjects* showed no significant polycythemia. The results suggest that "under the influence of adrenaline corpuscles equally saturated with CO , as those in the peripheral blood, are suddenly added to the circulation, thereby producing a true polycythemia without appreciable change in blood vol." L. A. M.

The site of action of thallium. HSI CHUN CHANG *Chinese J. Physiol.* 5, 70-82 (1931).—The depolarizing action of thallium is not dependent upon either thyroid or ovarian activity. Suggestive evidence is also furnished that nerve action is not involved. L. A. MAYNARD.

The effects of histamine, adrenaline and atropine. The sympathetic and vagus systems, acid-base equilibrium, movements of chlorides and water. P. SARACCHIVILI AND R. RAFFLIN *J. physiol. path. gen.* 27, 795-814 (1930).—The subjects used were epileptics and chronic alcoholics. Subcutaneous injection of histamine causes a reddening of the face, augmentation of pulse rate, increase in pH of the blood and a decrease in blood chlorides. Similar but less consistent results are obtained with adrenaline and atropine. It is concluded that the action of the neuro-vegetative system on the acid base equil. consists in the modification of the tissues for certain ions, such as Cl , which are retained or expelled according to whether the sympathetic or the vagus system is affected. L. A. MAYNARD.

Insulin and the hypoglycemic syndrome. V. ZAGAMI *Arch. fisiol.* 25, 339-71 (1930), *Physiol. Abstracts* 15, 593.—Insulin in direct application on the centers of the brain, has no influence on the nervous function. The influence on the blood sugar is the same as in subcutaneous application. A. E. MEYER.

Choline action on the cardiovascular system. VITTORIO SUSANNA *Arch. fisiol.* 28, 432-42 (1930).—Choline has a peripheral vasoconstrictor action. It causes a transitory decrease of the arterial pressure. In vagotomized and atropinized animals, a const. increase of the blood pressure is observed. The modifying effect on the heart function is due to an action on the sp. intracardiac app. A. E. MEYER.

Delayed absorption of pharmaceuticals caused by the vasomotor principle of the hypophysis (Vasopressin). G. SPAGNOL *Rev. sud-americana endocrinol. inmunol. ginecol. terap.* 14, 1-16 (1931).—In subcutaneous injections of the mixed solns. of vasopressin and strychnine or novocaine the absorption, as detd. by the onset of toxic symptoms, is considerably retarded. Vasopressin is still effective in a diln. of 1:15000000, whereas adrenaline has its limit at 1:100000. The action of vasopressin can not be explained by the pressor activity, but by a sp. influence on the permeability of the capillaries. A. E. MEYER.

Sodium bromide and thiosulfite in the treatment of eczema. PABLO LAVERZIO

Influence of radiation on the behavior of silver in the organism. I. Investigations on the whole organism of young rats. LUDWIG PRINCUSSEN AND WADIM ROMAN. *Biochem Z* 232, 202-8(1931)—In untreated young rats receiving injections of Ag_2SO_4 equiv to 0.6 mg Ag, 54% of the Ag is found in the organism in the form of AgCl , 28.6% in combination with protein and 17.3% as metallic Ag. Irradiation with the Vitalux lamp increases the portion bound to the protein to 57.2%, while it causes only a slight diminution in the metallic Ag fraction. On the contrary, radiation with the Hg lamp increases the Ag salt fraction to 61.2%, the other 2 fractions being somewhat diminished.

S MOROGLIS

The life-saving action of portions of plants and of the juices removed from them in the case of otherwise fatal subacute uranium intoxication. GREGG EISNER. *Biochem Z* 232, 218-28(1931)—Rabbits poisoned with very small amts of U almost invariably die when kept on a diet of oats and water but survive when to this are added beets or fresh green stuffs, also cabbage leaves. The fresh juice of the latter or boiled and filtered press juice likewise exerts a life-saving action. Aq exts made from alc.-treated material are also quite effective.

S MOROGLIS

Studies on the relationship between chemical composition and biological action. II. TAKAOKI SASAKI AND HIDENOSUKE UEDA. *Biochem Z* 232, 260-8(1931), cf C. A. 25, 1569—The Et esters of 2,3-, 2,4-, 2,5-, 3,4- and 3,5-diaminobenzoic acid have been prepd. The HCl salts of these were employed in animal expts. The 2,4-compd. was the most toxic for the mouse, the lethal dose being 3 mg in aq soln. (1:50) per 10 g. the 2,5-compd. is only half as toxic, while the other 3 isomers are about equally active and again half as toxic as the last. For the guinea pig the 3,4- and 3,5-compds. are fatal even in a dose of 0.01 g. per 100 g., while the 2,4- and 2,5-compds. are much less toxic than for the mouse, and the 2,3-isomer is the least toxic. Guinea pigs can at times survive even a subcutaneous injection of 0.1 g. per 100 g. These substances are all amyotatic poisons producing muscular motor incoordination with rigidity.

S MOROGLIS

Influence of adrenaline on the blood and organ lipoids. IRVING H. PAGE AND LYDIA PASTERNAK. *Biochem Z* 232, 295-309(1931)—After injecting 0.3 mg adrenaline every 15 min. for 4 hrs. into rabbits there was found a diminution in the phosphatides, cholesterol, fatty acids and total fat of the serum, while in the liver there was an increase in fatty acids, cholesterol and total fat. In the kidneys the phosphatides, fatty acids and total fat were diminished. In the brain the phosphatides were diminished, while the cholesterol was increased, and the fatty acids became somewhat less unsatd. In the heart also the phosphatides and cholesterol were diminished. The effect of adrenaline is the same on lipemic rabbits, though this may be marked by the progressing lipemia. Insulin and adrenaline both cause a lowering of the lipoids in the serum, but the effect of the former is somewhat more pronounced. So far as the influence on the lipoids of organs is concerned, with very few exceptions the two hormones do not act in the same way.

S MOROGLIS

Influence of arsenic and antimony compounds on the enzymic functions of the organism. VII. Buffer capacity of arsenous and arsenic salts. A. N. ADOVA AND I. A. SMORODNITSKY. *Biochem Z* 232, 459-68(1931), cf C. A. 25, 2193—The buffer capacity of a 1% Na_2HAsO_4 is about $1/4$ as great as that of NaH_2AsO_4 . Dilg. NaH_2AsO_4 with an equal vol. of H_2O causes a lowering of the p_{H} , diln. of NaH_2AsO_4 , if moderate, causes a shift of the p_{H} to the alk. side, but strong diln. (1:50) also causes a lowering of the p_{H} . The buffer effect of 1% Na_2HAsO_4 is the same as that of a phosphate buffer and is 6 times as great as that of the citrate buffer. To secure the same p_{H} in Na_2HAsO_4 and NaH_2AsO_4 solns., the normality of the former must be diminished 30 times. H_3AsO_4 has twice the buffer capacity of a phosphate and 12 times that of a citrate buffer toward alkali. The buffer capacity of 1% Na_2HAsO_4 is half as great as that of phosphate toward acid and only $1/4$ toward alkali. The buffer capacity of the phosphate mixt. for alkali is 6 times that of citrate, but for acid only $1/4$ times as great. The injurious effect of Na_2HAsO_4 injections is attributed to the disturbance of the acid base balance of the organism.

S MOROGLIS

The fate of colloidal iron administered intravenously. II. Long experiments. CYRIL J. POLSON. *J. Path. Bact.* 32, 247-69(1929)—Fe held by the lungs is essentially a foreign substance and is transferred principally to the liver. It is suggested that the path of the Fe from the lung is by way of the spleen. Fe is probably excreted by the cecum and the kidney, while part of the liver excess of Fe is transferred to the lymph nodes and remains in the body for some time.

JOHN T. MYERS

The antagonism between trypsin and insulin. ZDZISLAW ŚWIDER AND JULIAN WALAWSKI. *Med. Doświadczalna i Społeczna* 13, 1-24(1931), (In Russian, French

Abstract) —Insulin (I) lowers the tone of isolated rabbit intestine. Trypsin (II) restores the movements arrested by I. I and II when mixed have no effect. Boiled I arrests the intestinal movements, but boiled II does not restore them. A mixt of I and II boiled arrests the movements. I arrests the digestion of casein by II, this action being independent of the pH . It is possible that the antagonism between I and II is due to a combination of the two rather than to the destructive action of II upon I.

T. H. RIDGE

The use of therapeutic effects as end points in the biologic titration of the digitalis bodies. HARRY GOLD, BEN GELFAND AND WILLIAM HITZIG. *J Pharmacol* 41, 89-102 (1931) —Electrocardiograms were taken on cats after administration of tincture of digitalis (German leaf), tincture of digitalis (American leaf), tincture of digitalis (C11Cl₂ fraction), and tinctures of adonis, ouabain and digitoxin. The percentages of the fatal doses producing the following effects were noted: first change in T wave, or R-T interval, first increase in P-R interval, first appearance of ectopic beats and ventricular fibrillation. The degree of individual variability is least for the doses necessary to produce ventricular fibrillation, and by using the latter as the end point it is possible to standardize digitalis with greater precision than by the use of other phenomena. C. RIEGEL

Acquired tolerance to and cross tolerance between the nitrous and nitric acid esters and sodium nitrite in man. L. A. CRANDALL, JR., C. D. LEAKE, A. S. LOEVENHAART AND C. W. MUFFELBRAGA. *J Pharmacol* 41, 103-109 (1931) —Tolerance in human subjects (as measured by headache production) was established by periodic administration of erythrol tetranitrate, ethylene glycol dinitrate, Me nitrate and Am nitrite. No tolerance to Na nitrite was developed. Cross tolerance was demonstrated, and tolerance to Na nitrite was conferred by tolerance to the other substances. The effect lasted for approx. 10 days after discontinuation of the administration. There was no cross tolerance to histamine. Tolerance to the action on blood pressure and pulse rate was less easily produced and was less complete. C. RIEGEL

The acute toxicity of glyceryl trinitrate and sodium nitrite in rabbits. THEODOR V. OLTMAN AND LATHAN A. CRANDALL, JR. *J Pharmacol* 41, 121-6 (1931) —The lethal dose for sodium nitrite was 80-90 mg. and for glycerol trinitrate 45 mg. per kg. The symptoms of poisoning are given. C. RIEGEL

Guanidine structure and hypoglycemia. H. FAITZ BISCHOFF AND M. LOUISA LONG. *J Pharmacol* 41, 127-37 (1931) —The toxicity, effect on blood urea N, amino acid N and sugar and damage to liver and kidney are reported for di-*p*-phenylguanidine, di-*o*-anisylguanidine, benzimidazoleguanidine, benzothiazoleguanidine and benzoxazoleguanidine. C. RIEGEL

Further observations on the pregnancy response of the uterus of the cat. R. G. GUSTAVSON AND H. B. VAN DYKE. *J Pharmacol* 41, 139-46 (1931) —Subsequent to a series of 6-25 injections of urine obtained from pregnant women, the uteri of ovariectomized cats relaxed in response to epinephrine injections, while the uteri of animals having one or both ovaries intact contracted if the ovaries contained mature corpora lutea, or relaxed if no corpora lutea were present. C. RIEGEL

A study of the skin vessels in some forms of inflammation of the skin. G. H. PERCIVAL AND C. M. SCOTT. *J Pharmacol* 41, 147-63 (1931) —Histamine caused further dilatation, while adrenaline and Bier's spots overcame the existing vasodilatation of dilated minute skin vessels in exfoliative dermatitis, ultra violet erythema, oil of mustard erythema, psoriasis, dermatitis venenata and tinea corporis. The results fail to support the hypothesis that the vasodilatation has been produced by liberation of histamine, or a histamine like substance. C. RIEGEL

Intestinal bleeding following administration of posterior pituitary extracts. S. J. WEINBERG. *J Pharmacol* 41, 165-72 (1931) —Intravenous injection of toxic quantities of posterior pituitary extracts and of pitressin into normal unanesthetized dogs was followed in 6 of 14 cases by intestinal bleeding. C. RIEGEL

Chloroform poisoning during narcotization. A. SARTORI. *Chem.-Ztg.* 55, 222 (1931) —Qual analysis of the vital organs of an 11-months old child that died while subjected to C11Cl₃ anesthesia showed the presence of C11Cl₃, especially in the stomach and kidneys. Neg. quant. results were obtained, however. Hofmann's carbylamine reaction was used for the detection, the detn. was made by conversion of C11Cl₃ to HCl, and absorption in AgNO₃ soln. W. GORDON ROSE

The prevention of malaria with Schweinfurth green. A. MISSIROLI. *Rivista malarial* 9, 667-705 (1930)

Insulin therapy in malaria. C. TOSCANO. *Rivista malarial* 9, 734-40 (1930)

R. SANSONE

Plasmochin in chronic malaria. M. FRUTIERO *Rivista malaria* 9, 754-8 (1930) —Favorable results were obtained in 4 cases. R. SANSONE

Chemical constitution and physiological action. Behavior of the stereoisomers of α -bromoisocapronyl-L-asparagine. MADDALENA FURIA *Boll. chim. farm.* 70, 137-9 (1931), cf. C. A. 24, 829 —Differences were found in the physicochem. and physiol. properties of the stereoisomers of α -bromoisocapronyl-L-asparagine. The soly. of the l form (I) in H_2O was 1.8%, the content in the aq. soln. after extn. with oil was 0.503%, and the soly. in oil was 1.297%. The resp. figures for the d form (II) were 2.414, 0.688 and 1.720%. II produced sleep in shorter time in *Carassius auratus* L. than did I. When returned to fresh water, the fishes subjected to I regained their vitality without passing through a stage of excitement, in contrast to the fishes treated with II. G. S.

The influence of the combination of the alkaline earth chlorides with sodium chlorides on urethan hemolysis. R. MOGNAICKI *Protoplasma* 8, 215-21 (1929) —When the ion of Ca, Ba, Sr or Mg and the Na ion were present in about equal concns., the max. hindering effect on urethan hemolysis was observed. M. H. SOTTE

The action of narcotics of the urethan series on the colloidal activity of serum. P. J. JURISIC *Protoplasma* 8, 358-58 (1929) —The addn. of certain concns. of methyl and phenyl urethan caused gel formation when added to beef serum. The gelation was due to an alteration in the structure of the serum. M. H. SOTTE

The action of preparations of beer yeast on chemical processes in the liver and muscles of exercising animals. L. C. PI-SUREA BAYO, G. LISS AND T. OSYKA *Anales soc. esp. fis. quim.* 29, 117-9 (1931) —Previously (C. A. 25, 744) the effect of a prepn. from yeast autolyate on the chem. properties of muscle and liver of normal rats in exercise has been shown. These tests were repeated with pure, dry yeast (*Leuconose Blues*) on 20 white rats exercising 1 hr. daily and receiving a regular diet. Ten of these were fed 0.2 g. dry yeast in 2 cc. H_2O per day, the other 10 were held as controls. Glycogen, lactic acid and phosphoric acid content of liver and muscles were detd. after killing. Liver glycogen rose appreciably, muscle glycogen less. Lactic acid and phosphoric acid contents were slightly higher in muscle and liver of yeast fed rats. L. C. PI-SUREA BAYO AND G. LISS. *Ibid.* 29, 200-3 (1931) —Muscular effort was regarded as important as regular muscular activity. Dogs were fed daily 2 g. yeast per kg. of body wt. Controls had a normal diet. The exercise was 0.75-1 hr. daily on a treadmill, 48 hrs. in all. After quick killing the glycogen and lactic acid of liver and muscle and the fat content were detd. In yeast fed dogs the fat content was 6-8 times greater, the glycogen unchanged. Glycogen collected in the first few hrs. in liver and muscle is converted into fat, this proving the conversion of carbohydrates into fat in the organism. E. M. SIMES

The active principle of Hsi-sun. T. KONDO *Keijo J. Med.* 1, 223-47 (1930) —By distillation of the Chinese drug, Hsi sun, a yellow, slightly sticky oil was obtained. Large doses of this oil administered to frogs, rats and rabbits causes at first an increased irritability with a gradual central paralysis and finally a cessation of respiration. With small doses administered over longer periods of time, paralysis affecting first the extremities is the chief symptom. The conclusion is that this oil is the active principle of Hsi sun. A. J. VORWALD

The pharmacology of Panax ginseng. N. SUGIHARA AND P. MEN *Keijo J. Med.* 1, 347-408 (1930) —Panax when fed to rats augments the action of injected picrotoxin, camphor, caffeine and sodium cacodylate, but is antagonistic to the action of chloral hydrate, urethan, alc. and veronal. A. J. VORWALD

The pharmacology of Panax ginseng. N. SUGIHARA AND P. MEN *Keijo J. Med.* 1, 685-702 (1930) —Panax when fed to rats is antagonistic to the action of atropine, and adrenaline but is synergistic with pilocarpine and physostigmine. A. J. V.

Pharmacology of the active principles of ginseng. N. SUGIHARA AND K. KIM *Keijo J. Med.* 1, 711-72 (1930) —Panax acid, when administered to frogs, mice and rabbits, causes first stimulation of the spontaneous movements, then a gradual central paralysis. When applied directly to an isolated frog heart, it causes first a stimulation, then a paralysis. Panax acid dilates and then contracts the blood vessels. It brings about a decrease and final cessation of peristaltic movements of the bowel. Panacens, another active principle, acts likewise. A. J. VORWALD

Amino alcohols. VI. Preparation and pharmacodynamic activity of four isomeric phenylpropylamines. WALTER H. HARTUNG AND JAMES C. MURPHY *J. Am. Chem. Soc.* 53, 1675-9 (1931), cf. C. A. 24, 5578 —Four isomeric phenylpropylamines ($Ph-CH(NH_2)Et$, $PhCH_2CH(NH_2)Me$, $Ph(CH_2)_2NH_2$, and $PhCH(Me)CH_2NH_2$) were prepd. by catalytic hydrogenation of an appropriate intermediate. A preliminary pharmacol. examination of these compds. indicates that the optimum skeleton for pressor compds.

is $\text{PhCH}_2\text{CH}_2\text{NH}_2$, a shift in the relative positions of the Ph and NH_2 groups very greatly decreases pressor potency. Substitution of a Me group on either of the 2C atoms in the side chain of this skeleton confers oral activity, the presence of the secondary alc HO in $\text{PhCH}(\text{OH})\text{CH}(\text{NH}_2)\text{Me}$ serves to decrease the toxicity to a degree that becomes significant therapeutically.

C J WEST

Biological investigation of the active constituents of digitalis leaves. YOSHINO KOBAYASHI *Proc Imp Acad Tokyo* 7, 76-9(1931).—The LiOH ext was divided into 29 fractions by the use of various precipitants and solvents and their biol action studied. The glucosides easily sol in H_2O , which are not pptd by NaCl and $(\text{NH}_4)_2\text{SO}_4$, are less active than those insol or difficultly sol in H_2O , as far as cardiac action is concerned.

C J WEST

Parathormone dosage and serum calcium and phosphorus in experimental chronic hyperparathyroidism leading to ostitis fibrosa. AARON BODANSKY AND HENRY L JAFFE *Proc Soc Exp Biol Med* 27, 797(1930), *J Exp Med* 53, 591-604(1931).—On a low Ca intake hypercalcemia tended to disappear in chronic hyperparathyroidism on a given dose of parathormone (as large as 6 units per kg), apparently because of the reduction of a readily available Ca reserve. An increase of either the Ca intake or of the daily dose of parathormone caused a rise in serum Ca and symptoms of overdosage. Hypocalcemia developed in chronic hyperparathyroidism in young puppies on a low Ca diet. Tetany occurred at a Ca level which was higher and a P level which was lower than in tetania parathyreopriva of young puppies. About 0.1 g of Ca daily was apparently sufficient to maintain the serum Ca at a normal level. The serum P in chronic hyperparathyroidism in young puppies continued at or rose above the high level normal for young animals. Toward the end of long periods of treatment on large parathormone dosage (about 5 units per kg) serum P approached normal levels, pronounced hypercalcemia was absent but hypotonia and other symptoms of hyperparathyroidism were present. A single dose of parathormone caused early in the treatment and on liberal Ca intakes a more marked relative rise of serum Ca than in normal adult dogs, confirming previous observations. Later in the treatment and on low Ca intakes this effect was greatly reduced. Serum P rose after a single injection of parathormone, even when the effect on the serum Ca was slight or absent. The continued effect of parathormone on serum Ca after prolonged periods of treatment and the modified response of the serum P indicate tolerance due to some compensation, rather than immunity. The bone lesions, presenting the essential features of ostitis fibrosa cystica in varying degrees of severity, depending on the relation of the parathormone dose to the Ca intake and to the duration of the treatment, were most prominent on low Ca intakes, which permitted the use of large doses of parathormone without fatal hypercalcemia and without symptoms of overdosage.

C J WEST

The effect of iron on blood formation as influenced by changing the acidity of the gastroduodenal contents in certain cases of anemia. STACY R METTIER AND GEORGE R MINOR *Am J Med Sci* 181, 25-35(1931).—Responses of reticulocytes were observed (1) after small daily doses of Fe citrates with an alk. beefsteak meal, (2) after the same doses with an acid meal and (3) after the same doses followed by a 4- to 12-fold increased dose of Fe. The responses after feeding meals with high p_H were slightly less than after low p_H , the dose of Fe being const. The responses induced by increased doses of Fe indicate that the small doses of Fe were not optimal.

R C WILLSON

Dihydranol. Control of intestinal putrefaction in man by oral administration of 2,4-dihydroxyphenylheptane. VEADER LEONARD AND WM A FEIRER. *Bull. Johns Hopkins Hosp* 48, 25-38(1931).—The administration of dihydranol in doses of 0.3-0.45 g, thrice daily, destroys the true putrefactive flora of the intestinal tract with great certainty and regularity.

R C WILLSON

Thymophysin Temesváry. ERWIN E NELSON *J. Am Med Assoc* 96, 352-6(1931).—Experimentally no difference could be found in the pressor or oxytocic activity of pituitary alone as compared with pituitary plus thymus exts. The strength of "thymophysin" ampules examd by pressor and oxytocic methods was found to be from 25 to 33% of that claimed on the labels. When equiv doses of pituitary ext and "thymophysin" were compared on the excised uteri or on blood pressure, no differences could be demonstrated.

R C WILLSON

The renal changes following the intravenous injection of hypertonic solution of sucrose. H. F. HELMHOLTZ *Proc Staff Meetings Mayo Clinic* 6, 124(1931).—The epithelium of the convoluted tubules is specifically affected. Vacuolar degeneration is evident after 3 hrs. Within 48 hrs the swelling of the tubular epithelium is so marked that the lumen of the tubules is practically obliterated. The large cells, filled with clear protoplasm, resemble the cell of hypernephroma. The epithelium regains its

normal appearance within 10 to 15 hrs after the height of the swelling has been reached. As many as 22 injections at intervals of 5 days did not permanently reduce the output of phenolsulfonephthalein or increase the blood urea (1 exception). The chronic changes consisted of irregularities in the lining of the tubules, atrophy of the convoluted tubules lying close to the corticomedullary junction and increase in the connective tissue in this region. The glomeruli were unchanged. R. C. WILLSON

Synthesis and pharmacodynamic action of homoisopapaverine (and laudanosine) (KONEK) 10. Existence of 2 thioimidazole group in insulin (R'IZ, *et al*) 10. Hygienic importance of S. of sullured waters and S baths (GACNEWALD) 14.

Handbuch der experimentellen Pharmakologie. Band III, Hälfte 2. Edited by A. HEFFTER AND W. HEUBNER. Berlin. J. Springer. Cl C A 23, 3513.

Pathologische Anatomie und Histologie der Vergiftungen. Edited by ELSE PETAT. Berlin. J. Springer. 724 pp. M 144, bound, M 149.

12—FOODS

F. C. BLANCH AND H. A. LEPPER

The influence of pasteurization upon the preservation of buttermilk. LES FIK. *Chem. Listy* 24, 396-8(1930).—Sweet cream centrifuged from raw milk was churned for 3 hrs, yielding butter and buttermilk contg 0.9% butter fat. Aliquots (200 cc.) were heated to 87° and poured into wide-mouthed bottles. Unheated controls were also prep'd. All series were inoculated with spontaneously soured milk and were kept at 33°. A few were protected from the air by a layer of paraffin oil. The pasteurized buttermilk showed a lesser degree of lysis. Analyses for pasteurized buttermilk show total N 460 mg, N in casein 23, N in whey 176; analyses for untreated buttermilk show total N 460 mg, N in casein 15, N in whey 184. The heating made the protein more resistant to the action of bacteria. The lysis is anaerobic and occurs from 3 to 5 mm below the surface of the soln. The aerobic process occurs in the superior layer and leads to an oxidation of butter. The proteolytic bacteria cannot attack the proteins because of (1) the increased viscosity and coagulation of the protein and (2) the increased viscosity of the substrate and soln. FRANK MARESH

Chemistry in food freezing—storage. J. C. LAW. *Refrigerating Eng.* 21, 348-9 (1931).—Recent chem., phys. and biol. investigations on frozen and stored foods are discussed. A. H. JOHNSON

Humidity and air circulation in cold storage. M. W. BAOWNE. *Refrigerating Eng.* 21, 27-8(1931).—B. calls attention to the importance of humidity and air circulation in refrigerated rooms where food products are stored. Proper atm. conditions operate to prevent loss in wt., the development of bad odors and wilting and to assist in the maintenance of unvarying temp. and in the successful storage of the products for extended periods. A. H. JOHNSON

The composition of the wheat grown in Lettland. E. ZARINŠ AND O. AIZINA. *Acta Univ. Lettensis Kim. Fakultat. Series* 1, No. 4-6, Fasc. 2, Burtnica (in German) 123, (in Lettish) 127-37(1929).—The results of a study of the compos. of 36 samples of wheat are given in 3 tables. Table 1 lists the various amts. of water, crude protein, fat, starch, pentosans, sugar before inversion, sugar after inversion, crude fiber, mineral matter, Fe, Mn, CaO, MgO, H₂SiO₄, H₃PO₄, alkyl of the ash toward methyl orange, total alkyl, degree of acidity, p_{H_2O} , soly and sp. gr. Table 2 gives the detn. of catalase in the wheat. Table 3 shows the amylase content. RUSSELL C. EAB

Chenopodium quinoa Willd. C. M. ALBIZZATI AND R. FAURA. *Anal. soc. cient. Argentina* 111, 137-46(1931).—A comparison was made of the protein, fat, carbohydrate and mineral content of the quinoa seed with that of other cereals. Phosphorus and K predominate in the mineral fraction. A good white flour can be obtained from the seed, which can be mixed with wheat flour for the prep'n of bread. M. H. SOLLE

The F. A. Q. wheat sample. W. R. JEWELL. *J. Dept. Agr. Victoria* 29, 134 (1931).—Analyses of the annual sample of wheat of fair av. quality and of the flour, bran and pollard produced therefrom are tabulated. K. D. JACOB

Brabender apparatus for the fumigation of meal moths. JOSEF ŠPUSTA. *Chem. Obsor* 6, 59-62(62 English)(1931). J. KLČERA

A practical method for sorting flours according to their bran-meal contents. P. NOTIN AND M. LEMOIGNE. *Compt. rend. acad. agr. France* 17, 239-41(1931).—A soln. of pyrocatechol is added to a cup of flour. The flour is thoroughly mixed to obtain a homogeneous and not too liquid pat. At the end of 1/2 hr. the flours rich in bran are

rose colored, while those with less bran are much lighter in color. Phenol, guaiacol, hydroquinone, pyrogallol, gallic acid, diamphenol and tyrosine were also tested, and while they all gave good results, pyrocatechol was found the most satisfactory. The proportions of flour and pyrocatechol can vary between wide limits without modifying the results. While this method is not extremely accurate, it furnishes a simple and practical technic for classifying flours. J R ADAMS

The role of milk constituents in bread making. L A ALLEN AND J BELL. *J Roy Tech Coll Glasgow* 2, 550-63(1931).—A study has been made of the effect on bread of whole milk, sepd milk, whey and whey, from which albumin has been removed. All milk products improved the color, texture and flavor. Albumin has an inhibitory effect on loaf vol. This was shown by the fact that loaves made with whey had lower vol than the control, while loaves made with whey from which the albumin was removed had increased vol. Addn of egg albumin decreased vol, while milk which had been heated to coagulate the albumin increased vol. Dried sepd milk reduced loaf vol unless it had been boiled before drying. Lactose had no effect on vol, while lactic acid improved it. Whey products increased the rate of gas evolution, while milk reduced it. This indicates an inhibitory effect of the casein. AMY LE VESCONTE

The metals in the milk industry. POPP. *Landw Vers Sta* 111, 271-91(1931).—The order of decreasing resistance to corrosion of the 5 pure metals which were used is Ni, Cu, Fe, Zn and Al. The best alloy is a mixt of Fe, Ni and Cu. J R H

Testing milk. P GUARNIERI. *Atti III congresso naz chim pura applicata* 1930, 484-5.—A cryoscopic method for testing the water content of milk is advocated. E GEBAUER FUELNEGG

The freezing point of milk and its applications. G D ELDON AND J R STUBBS. *J Soc Chem Ind* 50, 135-41T(1931). cf *C A* 24, 4338.—A review and critical discussion are given of the methods in use for the detn of the f p of milk, especially in its application for the detection of added water. After a fairly extensive experience of the test, particularly of the Hortvet app, the authors are satisfied that it is of the highest value, not only for the detection of added water, but also for proving the purity of a genuine milk when the solids-not fat fall below the min allowed. An extensive list of references is given. J C JURJENS

The occurrence of peptidases in milk. ALFRED HEIDUSCHKA AND ERNST KOMM. *Z physiol Chem* 196, 187-94(1931).—Samples of cow, goat and human milk obtained under aseptic conditions were treated with glycyltryptophan and tested for free tryptophan by the Br reaction. The results were neg, indicating the absence of dipeptidase. Milk contg large nos of bacteria gave a pos reaction. Peptidase in milk is therefore of bacterial origin. A W DOX

Easy and exact method for ascertaining the presence of foreign fat in milk and butter. S CAMILLA. *Giorn farm chim* 80, 10-2, 15-8, 21-3(1931).—After describing the methods given in a previous paper (*C A* 24, 4339) C proposes another simpler method for the detn of foreign fat in butter. A sample of the butter to be examd is melted, and 0.3 cc of it is transferred to the butyrometer. Then are added 10 cc water-free ether, 10 cc 95% EtOH and after some agitation, 10 cc distd H₂O. The butyrometer is stoppered, well shaken and placed in a water bath at 37-40° until the fat has ascended to the top. Then the app is allowed to stand at 18-19° for 15-20 hrs. When the ethereal layer is limpid, the butter is probably not adulterated. However, when partial solidification of the fat occurs, an addn of foreign fat must be suspected. In order to ascertain the foreign nature of the solidified fat, the latter is examd under the microscope. The presence of typical stearin crystals is a pos proof of adulteration, since the foreign fats commonly used for sophistication (margarine, etc) contain stearin in various quantities. If the presence of coconut butter is suspected, the method just described cannot be used, since coconut butter contains little stearin, if any. In such a case 5-6 cc of the melted butter is emulsified with 100 cc of skimmed milk, and the presence of foreign fat is established by means of Quesneville's method. G S

Dissolving effect of sour milk on lead plates. J CSISZAR. *Kísérlet Közlemények* 32, 495-501(1929).—Soln of Pb depends on the acid no of the milk, time of exposure and quality of the Pb plate surface. In each expt a 0.1 mm thin Pb plate of 12 sq cm surface was treated with sour milk in a 100 cc Erlenmeyer bottle. Pb plates with metallic luster were not attached by milk of 24.5 Soxhlet-Hinckel acidity no on exposure during 26 hrs. Pb plates of a shiny but rough surface showed a decrease of 0.04 to 0.7 mg. Plates with an oxidized surface did not lose more than 0.4 mg Pb. Oxidized Pb plates with a rough surface were mostly attacked. Milk of 23.8 Soxhlet-Hinckel acidity dissolved 1.5 mg. Increase of acidity no also increased the quantity of dissolved Pb to a max. Then a decrease was observed. S S DE FINÁLY

Small-scale research on spray-drying of milk. A W. SCOTT *J. Roy Tech Coll. Glasgow* 2, 456-60(1931) —An app. with a drying chamber 6 ft. high and 1 ft. square is used to dry milk. Milk is forced under low pressure through a capillary tube with a nozzle of 0.005 in. diam. against compressed, heated air. From sepd. milk, a fine powder is obtained with a soly. of 99.5%.

The milk powders, single and mixed, and their hygienic control. VIRGILIO AARTERO *Boll. chim. farm.* 69, 937-43, 981-4, 987-9(1930) —A review. G. SCHWOB.

The utilization of dry skim milk in the manufacture of ice cream and cream cheese. J. C. MARQUARDT *N. Y. State Agr. Sta., Tech. Bull. No. 174*, 3-21(1931) —Chem. and phys. analyses of dry skim milk, including moisture, fat, titratable acidity and detns. of soly., were made at various intervals, and the quality of mfd. goods as related to the apparent soly., of the dry skim milk was followed throughout the work. It was concluded that dry skim milk can be kept at 40° or 70°F. in sealed containers for 8 months without any deleterious effects on its flavor or soly., but it was impossible to make good ice cream or cream cheese from dry skim milks that had been stored unsealed for more than 60 days. Structural changes were observed by the microscope. The H⁺ ion changes, however, could not be followed by colorimetric methods. J. K.

The effect of feeding raw and steamed potatoes on the yield and fat content of milk. K. RICHTER, K. E. FERBER AND N. ODAISEY *Miss. Arch. Landw. Abt. B. Tierernähr. Tierzucht* 4, 695-715(1931) —The yield and fat content of milk from cows on a normal diet were compared with those obtained when 12 kg. raw potatoes replaced an equal amt. (starch values) of the diet. The fat content did not vary, but the yield of milk decreased 5%. Potatoes were steamed and placed in a silo for 3 months, these potatoes were used in an expt. similar to the one using raw potatoes, the amt. of milk decreased 4%, and the fat content dropped from 3.4 to 3.14%. Analysis showed that the steam treatment had no appreciable effect on the chem. compn. of the potatoes. W. G. R.

Milk cows in nymphomania. F. MAYA AND K. WURSTER *Milchwirtschaft. Forsch.* 9, 335-8(1930) —The milk from 3 cows was studied. All had an unpleasant odor and taste. The fat varied from 3.87 to 4.9, total solids not fat from 8.0 to 16.2, Cl/lactose ratio from 1.49 to 1.80, % of the CaCl₂ serum from 41.4 to 42.1, elec. cond. from 37.7 to 42.1. Great care was exercised in selecting animals and collecting data, but the results are rendered of little value, because no data were obtained on the animals before the onset of this condition.

New chemical developed for testing butter fat. W. G. GOSS *Ice Cream Trade J.* 27, No. 5, 52(1931) —A modification of the Babcock test for butter fat is given. Accurate results are obtained in detg. the butter fat contents of ice cream, condensed milks, buttermilk, skim milk, cream and milk.

Determination of biacetyl in butter. G. TESTONI AND W. CIUSA *Ann. chim. applicata* 21, 147-50(1931) —As a possible method for identifying pure butter, a method has been developed for detg. the biacetyl in butter. The biacetyl is sepd. from butter by steam distn. and treated in boiling soln. with NH₂OH, and a few drops of a 0.1% NiSO₄ soln. and NH₂OH (conc'd). A red ppt. of nickel dimethylglyoxime is formed. The ppt. may be weighed or estd. colorimetrically. Melted and pasteurized butter as well as margarine do not contain any biacetyl, whereas fresh butter contains about 0.0005%. Its presence becomes a measure of the freshness of the sample of butter.

Fungi found in butter. M. GAINES, V. C. E. KENNELLY AND H. A. CUMMINS *Sci. Proc. Roy. Dublin Soc.* 19, 543-69(1933).

Manufacture of low-acid rennet cottage cheese. H. L. WILSON AND C. S. TRIMBLE *Creamery and Milk Plant Monthly* 20, No. 5, 27-32(1931) —The process of mfg. low acid rennet cottage cheese is described. When the milk contg. a starter and rennet has developed an acidity of 0.50 to 0.55%, it is cut into cubes. Water is then added and the curd is cooked and finally washed. Before creaming with cream or skim milk the curd is allowed to set in the refrigerator for 12 hr. At the time of creaming 1.5 lbs. of NaCl is added per 100 lbs. of dry curd.

Composition of some rabbit carcasses. W. K. WILSON *J. Ministry Agr.* 36, 1203-6(1930) —The chem. compn. of rabbit flesh (Angora) closely resembles that of chicken, the av. protein content differing by less than 1%. The compn. of the flesh from male and female rabbits was noticeably different, especially in the fat content, females contg. 4-6% more fat than the males.

Use of acetaldehyde in the storage of fruit. S. A. TROTT AND R. G. TOMKINS *J. Council Sci. Ind. Research* 4, No. 1, 6-11(1931) —Fruit stored in atm. contg. small quantities of AcH sometimes remained in sound condition for a longer period than fruits stored in air. AcH may be effectively introduced into fruit stores in sufficient concns.

to check wastage due to mold growth without injuring the fruit. The cones controlling mold development without injuring the fruit is about 1,500. Expts were made with grapes, strawberries, raspberries, cherries, plums and oranges. The Acell effectively controlled mold growth and decay in all fruits in cones varying from 1,250 to 1,000. Control was not obtained by dipping fruits in solutions of Acell. C. R. FELLERS

Ethylene oxide as a new fumigant for dried fruits. J. I. THOMAS. *J. Council Sci. Ind. Research* 4, No. 1, 53-4 (1931).—Limited tests showed that 2 lb. of ethylene oxide per 1000 cu. ft. in enclosed rooms was fully effective in destroying eggs, larvae and pupae of the dried fruit moth, *Plodia interpunctella*, in 56 lb. boxes of artificially infested sultana raisins. C. R. FELLERS

Studies of tomato quality. IV. Variability in quality and food value of tomatoes. JOHN H. MACGILLIVRAY. *Proc. Indiana Acad. Sci.* 38, 159-63 (1929).—In tomatoes taken from the same field dry matter varied from 6.0 to 10.62%, total N from 1.74 to 5.72%, acidity from 3.28 to 11.63% and reducing sugars from 32.38 to 45.70% of total solids. H. M. STARK

Studies of tomato quality. V. Clearing is not essential in determining reducing sugars of ripe tomato fruit extract. J. H. MACGILLIVRAY AND A. H. WATSON. *Proc. Am. Soc. Hort. Sci.* 26, 137 (1929), *Dept. Sci. Ind. Research Index to Literature of Food Investigation* 2, No. 2, 56 (1930). E. J. C.

Organic acids of spinach, broccoli and lettuce. E. K. NELSON AND H. H. MORTON. *J. Am. Chem. Soc.* 53, 1909-12 (1931).—Fresh spinach contains 0.31% (CO_2H)₁, citric acid and a small quantity of malic acid were separated by the ester distillation method. Analysis of broccoli shows that the leaves and buds have materially the same composition and nutritive value. Both buds and leaves contain proteins somewhat in excess of that reported in spinach. The predominating organic acid in broccoli is citric acid, it also contains L-malic acid and small quantities of oxalic and succinic acids. The proportion of citric and malic acids is 3:2. The organic acids of lettuce were found to be oxalic 0.011%, L-malic about 0.065%, and citric about 0.048%. C. J. WEST

Utilization and composition of oriental vegetables in Hawaii. H. L. CHUNG AND J. C. RIPPERTON. Hawaii Agr. Expt. Sta., Bull. No. 60, 64 pp. (1929).—The proximate composition, mineral elements and alkali of the ash were determined in 56 varieties of oriental vegetables commonly used for human food in Hawaii. Of the group of 30 leafy and stem vegetables, all are comparatively rich in Ca with the exception of bamboo shoots, small bean sprouts and tender fern fronds. The vegetables classed as Ca deficient are all immature stems. Large bean sprouts are exceptionally high in protein, fat and carbohydrates. The several cabbage-like plants are characterized by high water content, comparatively low protein and energy values, and high Ca. The superiority of the green leaf, non-heading Chinese cabbage over the blanched heading variety is marked. Kale, radish and turnip greens stand out as rich sources of Ca. Of the spinaches, the Chinese variety is superior in Ca but inferior in Fe. Both have a decided excess of basic elements. Honeysuckle and matrimony vine are rich in protein, energy constituents and minerals and merit greater utilization in the human diet. Matrimony vine contains nearly twice as much of the mineral constituents, except P, as does any vegetable in this group. Twelve fleshy or pod vegetables were examined, the former are characterized by high water content and low protein and energy values but are very succulent, and the latter by their high protein content and generally high minerals. The balsam pear is remarkably high in P (0.107%) and base-forming elements. The soy bean has not only a high protein content but the protein is of very high quality. Fourteen aquatic and starchy root vegetables showed generally high N free extract, low protein and Ca and moderate amounts of P. Kudzu and Chinese yam are among the most desirable foods in this group. The lotus root and water chestnut are important and nutritious foods. Bibliography. C. R. FELLERS

Determination of sand in vegetables. I. RÖR. *Chem.-Ztg.* 55, 221-2 (1931).—Introduce 5 g. of an air-dried sample into a Späth sediment glass, filled $\frac{1}{2}$ full with CCl_4 . Let stand for 1 hr., stirring several times with a wire. Pour off most of the CCl_4 , transfer the residue quantitatively to a porcelain dish, and remove the organic matter by heating at a dull red heat. For the determination of sand in rice, add 20 cc. 30% NaOH to the ash, heat on a water bath for 15 min., filter, wash and ash. W. GORDON ROSE

Decarboxylation studies on pectins and calcium pectates. C. M. CONRAD. *J. Am. Chem. Soc.* 53, 1909-2003 (1931).—Decarboxylation of apple and lemon pectins with mineral acids occurs at low concentrations of acid and is very appreciable with as weak as 0.1% H_2SO_4 . The extent of decarboxylation of apple pectin in various concentrations of H_2SO_4 , and HCl and of lemon pectin in various concentrations of H_2SO_4 is reported for a 16 hr. period of boiling. Decomposition of the galacturonic acid of apple pectin is complete with

12% HCl but is not complete with 18% H_2SO_4 . A comparative study of the Ca pectate from several sources shows that the rate of decarboxylation with 2% H_2SO_4 is approx the same in each

C. J. WEST

Photomicrographic studies of sucrose crystals SYBIL WOODRUFF AND HELEN VAN GILDER *J Phys Chem* 35, 1355-67 (1931) —Photomicrographs were made of fondants prepared by partially hydrolyzing sucrose by boiling with one half its vol of water and $\text{KHC}_4\text{H}_4\text{O}_6$, tartaric acid, citric acid or HCl and also of fondants prepared by boiling aq solns of sucrose and glucose or fructose. The moisture content of fondants was about 13%. Glucose, fructose and invert sugar served equally well in regulating the size of crystals. 6-15.8% of invert sugar or 7% glucose or fructose gave most desirable crystal sizes. Syrups contg 43.43% invert sugar or more would not crystallize, and those contg 16.3-23.56% formed a semifluid mass of crystals. Candy observed to be coarse in texture contained crystals 45μ in size, below 20μ seems desirable, as the sense of touch in the mouth can detect gritty texture when the crystals are 25μ in size but not when they are below 19.6. The difference of 6-10 μ represents the difference between a smooth and a gritty texture. Storing the fondant does not appreciably alter crystal size.

ODEN I. SHEPPARD

The determination of phosphoric acid in organic substances especially feed stuffs. W. LEPPER *Landw Vers Sta* 111, 159-61 (1930) —Five g of the substance and 2 g CaCO_3 are heated to ash. With P_2O_5 rich material (0.5% or more) the ash is washed into a 500 cc flask the dish rinsed with HNO_3 and the ash dissolved without heating by shaking with about 20 cc HNO_3 . The flask is filled to the mark with water, and the P_2O_5 detd in 15 cc of the filtrate by the method of Lorenz. With P_2O_5 poor material the method is the same, except that a 250-cc flask and 50 cc of the filtrate are used.

JOHN R. HILL

The use of copper sulfate in place of mercury for the Kjeldahl analysis. W. LEPPER *Landw Vers Sta* 111, 155-8 (1930) —For the detn of crude protein in feed-stuffs 1 g of material is heated with 5 g of $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$, 15 g K_2SO_4 and 20 cc H_2SO_4 until the soln has a clear green color (about $\frac{1}{4}$ hr). The soln is heated for $\frac{1}{2}$ hr longer and distd in the usual manner with NaOH.

JOHN R. HILL

An innovation in meat packing waste treatment (HALLORSON) 14. Recent investigations on the constitution of pectins (BRIDEL) 10. Organic foodstuffs and methods of food analysis (RUMER) 11E. The influence of frequency of cutting on the productivity botanical and chemical composition and the nutritive value of "natural" pastures in Southern Australia (DAVIES, SIM) 11D. Evaporating milk (U S pat 1,799,478) 1. Fats and waxes (Fr pat 697,572) 27. Irradiating foodstuffs (Austrian pat 121,657) 17. Dyes [for preserving foods from fermentation] (Fr pat 698,076) 25.

Carbohydrate foods ALOIS JODLBAUER and FRANZ KLEINVIKEL, Austrian 121,732, Oct 15 1930. The albumin content of carbohydrate foods, especially chocolate or cocoa or confectionery made therewith, is increased by incorporation of yeast.

Food material from malt-grain residues. MARGARET M. KOLLER, U S 1,799,142, March 31. In treating the residue collected after extg sol constituents from malt grain the acid pressed juice "is sepd from spent grain," and an elec current is passed through the juice for a sufficient time to effect a substantially complete pptn of the suspended solids in the juice and to render the juice alk in character, so that fermentation of the solids is inhibited. These solids are suitable for use as an animal food. App is described.

Increasing the baking capacity of flours. F. GRUZZI and A. SZABO, Hung 102,064, Jan 18 1930. Peroxides stabilized by addn of enzymes or org compds are mixed with the flour or with the leaven. E g 3 g asparagine is added to 10 cc H_2O_2 and cooled. The cryst product is mixed with 3 g malt diastase, and 0.1-0.5 g of the mixt is added to 1 kg flour or leaven.

Apparatus for pasteurizing and cooling milk. GEORG RÖTTGER, Ger 516,818, Aug 28 1928.

Apparatus for sterilizing milk by heating and cooling. BERGEDORFER EISENWERK A-G ASTRA WERKE, Ger 516,812, Nov 20, 1929.

Milk of high vitamin and low bacterial content. ROY R. GRAVES, U S 1,798,413, March 31. Milk is taken from the cow and transferred to a pasteurizing app under vacuum conditions without release of the vacuum during the transfer, and the milk is allowed to remain under vacuum conditions at body temp for 30-60 min before being pasteurized.

Sour milk product WALTER L. KULP, U S 1,799,303, April 7. There is added

to skimmed milk about 1% of concentrated tomato juice in which about half the titratable acidity has been neutralized, the mixt is stirred, heated at 66-69° for about 1-2 1/2 hrs., cooled to about 38°, and there is then added about 0.25% of a pure culture of *Lactobacillus acidophilus* developed in milk, and the mixt is held at a temp. of about 36-38° until a curd is formed. The product may also be made with use of yeast instead of tomato juice.

Tubular heat-exchange apparatus suitable for heating or cooling milk or cream. FRITZ G. CURNELL, JR. (to Jensen Creamery Machinery Co.) U. S. 1,799,356, April 7. Structural features.

Cheese. ALEXANDER AXELROD. Fr. 698,478, Dec. 19, 1929. Milk is cooled to 30-40°, and submitted to fermentation with ferments such as *Bacillus caucasicus*, kephir yeast, *Bacillus bulgaricus* or peptonizing bacteria, if necessary with the addition of agents to neutralize the lactic acid formed. The mass is then thickened to about 1/4 of its vol. treated with acid and pressed.

Device for drying and smoking meat, etc. E. WINKELMÜLLER & Co. Ger. 516,753 Mar. 15, 1927.

Use of refrigeration and carbon dioxide in preserving fruits and vegetables, etc. CHARLES BROOKS (dedicated to the free use of the public). U. S. 1,798,781, March 31. Food products to be preserved are chilled to a low temp. suitable for preservation and simultaneously submerged in a bath having a CO₂ content of 15-40% at the beginning of the cooling and which is decreased to 0-10% when a suitable preserving temp. is attained.

Conserving animal and vegetable foods. RICHARD WILLSTÄTTER. Ger. 516,923, Sept. 11, 1929. Addition to 513,115 (C. A. 25, 1921). In preserving food by HCN as described in 513,665, inert gases such as air may be present in the HCN current. Cf. C. A. 25, 2494.

Preserving fruit juices, etc. HENRIETTE A. LIESER NEH LANDAU. Fr. 698,297, July 2, 1930. Fruit juices, jams, etc., are preserved by bringing them in contact with sugar, either solid or in soln. and with a fungicide such as HCOOH, whereby, due to the osmotic action of the sugar and the action of the HCOOH, a sterile product is obtained.

Apparatus for freezing liquid confections, etc. HANS K. CHRISTENSEN and LOUIS J. HENNETT. U. S. 1,794,725, March 31. Various details are described of an app. comprising a cylindrical chamber with a screw conveyor within it and a pipe coil around it for circulation of cooling fluid.

Pectin. ROGER PAUL and ROBERT H. GRANDSEIGNIE. Fr. 698,161, June 28, 1930. An app. is described for treating pectic juice with acetone to obtain purified white pectin. Cf. C. A. 25, 1299, 2193.

Pectin solution from apple pomace. ELTON R. DARLING and HOWARD F. MACMILLIN (to Hydraulic Press Mfg. Co.). U. S. 1,799,140, March 31. The pomace is subjected to high pressure to remove any remainder of platable juice, it is then broken up, and not more than 10% by wt. of hot water is added, and the mixt. is subjected to much higher pressure to express the pectin from the pomace.

Beverages. EDUARD O. ORNFELDT and MAKOTO LOEW. Fr. 698,149, June 27, 1930. "Mate" or Paraguay tea made from the dried leaves of "*Ilex paraguayensis*" is improved in taste and smell by the action of carbonyl compounds, e. g., AcOH or ketones.

Working up lupine seeds. GLS. FÜR LUPINENINDUSTRIE M. N. H. Ger. 522,335, June 24, 1924. After the seeds have been steeped in known manner, the swollen seeds, which are those from which the poisonous constituents have been adequately removed, are sepd. from the unswollen or slightly swollen seeds by sifting. Cf. C. A. 25, 1298.

Preserving green fodder. GUSTAV FINGERLING. Ger. 522,333, Nov. 20, 1925. The fodder is steeped for about 8 days in water which is protected from the air by an oil layer and contains a bactericide in a concn. sufficient to disinfect the fodder without rendering it inedible. Alternatively, a stronger disinfectant soln. may be used, and replaced after a time by water. Cf. C. A. 25, 1923.

Method and apparatus for preserving green fodder by the cold-pressing process. ADOLPH HUPERTZ. Ger. 522,731, Jan. 8, 1926.

13—GENERAL INDUSTRIAL CHEMISTRY

HARLAN S. MINER

Acid-resisting materials for chemical and building industries. P. N. GRIGOR'EV and I. I. SILVISTROVICH. *Zhur Prikladnoi Khim.* 3, 1155-8 (1930). —Na₂SiF₆ is recommended as a binder for materials containing silicates. V. KALICHYNSKY.

The steam and electric power plant of Imperial Chemical Industries, Ltd., at Billingham H A HUMPHREY, D M BURST AND J. W. BANSALL *J. Inst. Elec. Eng.* 68, 1233-75, 1276-90(1930) E J C.

Blowing bulk chemicals about the plant. E H. DE CONINGH *Chem. Markets* 28, 497-499, 501, 503(1931) E H.

Safety in handling compressed gases. ROBERT H FERGUSON *Chem. Markets* 28, 505, 507, 509-511(1931). E H.

Smokes and smoke gases. E SMOLCZYK *Z. ges. Schiess.-Sprengstoffw.* 26, 132-5 (1931) —A discussion of the products of a fire from the standpoint of danger to exposed personnel particularly firemen. The greatest danger comes from the smoke, composed of both solid and liquid particles. These are present in the largest quantity early in the fire before the temp is sufficiently high to prevent the approach of firemen. They frequently contain substances which deaden the nerves in the mucous membrane and act as narcotics, causing sudden collapse. The mechanism of this action is described. Less dangerous are irritant gases and CO. Suitable gas masks contg smoke filters are described. A L KIBLER

The toxicity of certain benzene derivatives and related compounds. HENRY F. SMITH *J. Ind. Hyg.* 13, 87-96(1931) —The toxicity of many compounds used industrially has been studied. The blood changes which Selma Meyer (*C. A.* 22, 1417) found after benzene, toluene and xylene poisoning have not been in agreement with those reported by American investigators. Long exposure showed evidence of lung inflammation and degeneration of internal organs. Xylene and toluene, because of their higher b. pts., are not so hazardous as benzene, and they are not absorbed through the skin to any extent. Aromatic nitro compounds act on the central nervous system, whereas amino compounds are blood poisons. Headache, dizziness, visual disturbances, cyanosis, abdominal cramps, coma and convulsions are characteristic of both. Poisoning from amino derivs is less serious but they are absorbed readily through skin as well as through the lungs. Aniline poisoning may result in malignant bladder tumors, 25 g is a toxic dose, but 0.4-0.6 mg per l may be breathed for about 1 hr without much effect. Toluidines produce less cyanosis than anilines. The diamines used in dyeing furs may be very toxic, and poisoning is sudden. Some destruction of red cells results. A study of guanidine derivs revealed a sharp line between the fatal dose killing in a few hrs and the dose from which there was recovery. Death from these resembled cyanide poisoning. All these benzene poisons produced one outstanding lesion, namely a tendency to extravasation of blood in the lungs. This varied from scattered minute petechiae to massive lobular or lobar hemorrhage. The lesions were present in exptl animals and in man. In industry, accidental poisoning with any material having a lethal dose over 2.0 mg per kg per os is rare. With a limit of 100 mg per kg there is little danger if reasonable precautions are taken. Personal idiosyncrasy and the fact that absorption through the skin is often very rapid must be kept in mind. G H W. LUCAS

Destruction of mustard gas (dichlorethyl sulfide) in the ground by means of fire. HUGO STOLTZENBERG *Z. ges. Schiess.-Sprengstoffw.* 26, 135-7(1931) —An app is described. A L KIBLER

Commercial cooling units. F E STEWART *Refrigerating Eng.* 21, 21-6(1931) —The effect of the spacing of fins and design of the box interior are considered in relation to heat transfer characteristics. It was noted that max heat transfer was dependent on the total heat load, the resistance offered to free air circulation and the depth of the fin in the direction of air travel over the surface. The most effective spacing varied from 3/4 in for fins 4 in deep to 1 in for fins 14 in deep. A H JOHNSON

Refrigerant properties—some refinements. H D EDWARDS *Refrigerating Eng.* 21, 106-7(1931) —The phys. consts of 15 refrigerants are tabulated. A H J.

Economical marine insulation. E B JOELSON *Refrigerating Eng.* 21, 13-20 (1931) —Peculiarities in marine insulation, as well as insulants suitable for ship's holds, are discussed. Methods and formulas are given for computing economically practical insulation for ships. Charts are included to illustrate the use of the author's formulas. A H JOHNSON

Molding powders for electrical use—wide range available. J R. DOLPHIN *Times Trade and Eng. Suppl.* 28, No 664, 20(1931) —Illustrations of hot and cold molding are shown. F M SYMMES

Effect of surface conditions on heat transmission. S J KOHLI *Trans. Inst. Chem. Eng.* Advance copy, 19 pp., Oct., 1930 —The over all and film coeffs of heat transfer through a Cu tube, within which a current of hot air passed, and which was cooled with water outside, were studied with special reference to variations in the character of the surface. Results were obtained for the over all coeff in fair agreement

with Joule and Ser's formula, $k = 160$ to $17\sqrt{v}$, where \sqrt{v} is the gas velocity in m/sec. K's results are summarized by the equations k (over all) $= 16.912v^{0.48}$ and k (air-Cu) $= 16.116(w/a)^{0.48}$, where w/a is the mass velocity. Other formulas are discussed and criticized. Coatings with red, green and violet enamels increased the heat transfer throughout. It is suggested that radiations of short wave lengths may be absorbed and converted into heat energy. At high velocities convection plays an overwhelming part, and therefore the thickness of the air film is important. At lower velocities radiation and absorption are of more importance. The metal water film coeffs. cited in a few cases show an increase with the logarithmic mean temp. difference, undoubtedly because of increased convection. These coeffs. are markedly increased by black surfaces on the gas side. R C A

Porous masses from gypsum [for heat insulators] (Ger. pat. 522,139) 18.

BRANEGAN, JAMES A. Popular Industrial Chemistry; a Handbook for Manufacturers, Dyers, Tanners, Textile and all Industrial Workers. Philadelphia: Westbrook Publ. Co. 164 pp.

Chemical Engineering and Chemical Catalogue, 1931. 7th ed. Edited by D M. N. WITT. London: Leonard Hill, Ltd. 65s. Reviewed in *J Pharm J* 126, 386 (1931).

GRAHAM, C. J. Patent Rights for Scientific Discoveries. Indianapolis: The Boldt Merrill Co. 286 pp. \$5. Reviewed in *J Western Soc Eng* 36, 136 (1931).

OSSE, J. F. V. Warenkennzeichen Technologie. Amsterdam: D. B. Centen. 1200 pp. 1 men, 11 25 50.

FAH, FERDINAND, and ARNDT, WALTER. Die Rohstoffe des Tierreichs. Bg. 6. Berlin: Gebrüder Bornträger. Pp. 419-576. M 12. Cf. C A 21, 5899.

Schutz- und Angriffsrichtungen, Reaktionen auf Schädigungen. Edited by M. ASKANAZY, et al. Berlin: J. Springer. 897 pp. M 02, bound, M 07 80.

Gaseous reactions. NIKOLAI CARO and ALBERT R. FRANK. *Fr* 608,160, Aug. 17, 1929. In gaseous reactions substances are added to the gases which may afterwards be removed. Thus in the production of N_2O_4 or HNO_3 by the catalytic oxidation of NH_3 or HCN , steam is added to the gases before the reaction zone and is condensed after the reaction.

Separating carbon monoxide from gases. HAROLD KEMMER. *Fr* 607,752, June 21, 1930. CO is sep'd from gases contg. it by converting the greater part of it into CO_2 and H_2 by means of steam and catalysts and the remainder into CH_4 and water by hydrogenation. Details of the process and an app. are described.

Treating gases with liquids. ELMER DEPASSE. *Fr* 608,713, Oct. 14, 1929. The known method of washing a vapor by passing it through a liquid, with the use of venturi tubes, is extended in the neutralization, superheating, cooling and purifying of gases, vapors and fluids.

Recovering organic substances from aqueous solutions or emulsions, or from gas mixtures. ISHOR FRAUEN. Ger. 522,443, Mar. 8, 1928. The solids, mixts., etc., are brought into contact with a powder ore, e. g., galena or zinc blende. The ore adsorbs the org. substances, which are then recovered by treating the ore with a hydrophilic colloid, e. g., gelatin or albumin, or with an org. comp'd of stronger polarity than the adsorbed comp'd. Thus, CaH_2 may be sep'd from a mixt. with water by shaking the mixt. with powder galena, the adsorbed CaH_2 is liberated by addn. of a little oleic acid. Other examples are given also. The method may be applied to the purification of water.

Dispersions of solids. I. G. FARMING A. G. *Fr* 607,009, April 24, 1930. Stable dispersions of metals, metalloids or their mixts. or alloys are prep'd by disseminating the solids in a finely divided state with rubber or like masses previously masticated and dissolving the mixt. in a solvent after the dispersion is pushed to the point where a sample forms a soln. in an appropriate solvent without a deposit even on standing.

Recovery of solvents such as those used in the manufacture of the alkali metal phenylglycines. JUSTIN F. WARR (to National Aniline & Chemical Co.) U. S. 1,798,713, March 31. Aniline, an alkali metal cyanide and CH_3O are caused to react in a solvent comprising alc. and water to form a reaction mixt. contg. an alkali metal phenylglycine, alc., aniline, NH_3 and water. Liquids are vaporized from the reaction mixt., the vapors are condensed and NH_3 is sep'd from the condensate. Substantially all the condensate is refluxed and brought into contact with the vapors, and the operations of distn., condensation, sep'n. of NH_3 , and return of condensate are continued until the

condensate contains only a small quantity of NH_3 . A portion of the condensate is then withdrawn from the system, while the distill and condensation are continued. App is described.

Condensable solvent containing hydrocarbons. WILK KANTOROVSKY and ELLIOTT MERRILL. Fr 675,501, June 17, 1930. Light hydrocarbons used for cleaning, etc., are rendered condensable by a solvent composed of a light petroleum distillate and a compound of the formula $\text{R}_3\text{N} - \text{N} - \text{XFA}$, in which R is H, an alkyl group or a substituted alkyl group, X is an alkyl group or a substituted alkyl group, and FA is a fatty acid such as oleic, stearic or palmitic acid.

Desulfurizing extraction agents. GEORGE SCHMIDT. MARSHALL STUBBS. Ger 517,001, Apr. 20, 1932. Ammoniacal extraction agents used in desulfurizing materials are desulfurized by passing into an agent, such as Ca cement or dry iron ore which absorbs H_2S , but leaves the FeSO_4 still in the free state.

Desulfurizing caustoda. WALTER O. ROBERTSON to New Jersey Zinc Co. U S 1,704,577, April 7. For promoting the desulfurization of oil is in a liquid medium such as mineral pulp, the material is treated, without added heat, with propyl sulfide vapor.

Heat carriers for high temperatures. I. G. FARRINGTON A-G. Fr 675,313, June 13, 1930. A mixt. of cry. bodies stable at 300° , such as borax or borax mixts. of CaF_2 , PbO and Pb_2O_3 , are used as heat carriers.

Drying storage-battery plates or other articles with heated gases. WILLARD L. REYNOLDS to W. Ward Storage Battery Co. U S 1,709,245, April 7. The articles are placed in a chamber through which a heated gas is passed until a predet. amt. of the moisture has been carried away, and the chamber is then connected in series with another chamber containing wet articles. An arrangement of app. is described.

Filter for gas and smoke masks. HARRY A. KIRBY and Wm. A. BORLE (Borle to Kirby). U S 1,708,194, March 31. A fibrous cellulose material is impregnated with activated charcoal in a definite medium.

Brigs for refrigerators, etc. ROBERT MERRICK. Austrian 121,988, Nov. 13, 1932. The corrosive action of the brine on metals is reduced by adding a small quantity of an cry. nitro compound, e. g., picric acid.

Condenser with flat surfaces suitable for condensing gases in refrigerating systems. GEORGE HILDEY. U S 1,708,703, March 31. Structural features.

Condenser suitable for ammonia condensation in refrigeration. GRIGORI H. WITTE. U S 1,708,824, March 31.

Pipe-coil condenser suitable for condensing ammonia. GEORGE M. J. BLOMBERG to Platen-Merrers Refrigerating System, Alton, Ind. U S 1,709,081, March 31.

Insulating materials. I. G. FARRINGTON A-G. Fr 675,312, June 13, 1930. Insulating and coating materials are made by mixing and molding under heat and pressure waste wood, peat or residues from the acid treatment of peat, or lignin from the saccharification of wood, with condensation products of urea and aldehydes or their polymers. Cf. C. A. 24, 734.

Heat-insulating material suitable for lining floors, ovens, etc. ARTHUR J. REES (to Armstrong Cork Co.) U S 1,709,233, April 7. After laying down a substantially dry facing mixt. of construction material such as cement, sand and asbestos, there is placed over this a body material comprising a clay mass with earth, moist, water, and liquid from the body material is utilized by the facing material.

Waxing paper and wrapping it around wire. NATHAN C. LAMBERT (to National Electric Products Corp.) U S 1,708,889, March 31. App. and various details of operation are described.

Electrical insulators. ELMER O. BRUCE (to Brown Co.) U S 1,709,183, April 7. Impregnated insulators are formed with a stem member composed of long wood fiber and a compressed absorbent filler such as wood flour and with a flange member having a greater amount of absorbent filler and a smaller proportion of fiber.

Cold-molded insulation. ROY E. COLEMAN (to Moconway Elec. Corp.). U S 1,709,034, March 31. A binder is formed by heating and mixing an asphaltic substance, a drying oil and a dispersing medium in the form of a fatty acid or a drying oil. The binder is mixed with filler material such as finely divided asbestos, and the mixt. is cured.

14—WATER, SEWAGE AND SANITATION

EDWARD BARTOW

Recent progress in water supply. C A EMERSON, JR., H B CLEVELAND, H F. FERGUSON, J C PRITCHARD AND ROBERT S WESTON *Public Works* 61, No 6, 23-4 (1930) —The report discusses advances made in methods for the improvement of water color, taste, odor and softening. For color removal superchlorination followed by dechlorination is effective when the Fe content of the coloring matter is high. Chlorinated copperas has been used effectively, as well as filters loaded with ferric hydrate. Iron control and improved mixing devices have decreased the cost of color removal. Progress has been made in eliminating mud balls from filters. The use of air during filter washing has been recently introduced. Filter sand may be kept clean by employing velocities sufficient to produce 50-60% expansion of the sand bed. Variations in temp must be corrected for by varying the wash water velocity. Cleaning of filter sand may be accomplished by the use of caustic soda. Improved methods for recovering phenols at coking plants have reduced considerably the tastes and odors caused by industrial wastes. Activated C filters are especially effective for removing objectionable odors and tastes. Another method is pre ammoniation. Improvements in filter plants include the use of mech. devices to replace cleaning by hand and the utilization of the coagulating effect of sludge from the coagulating basin on the silt contained in raw water. The zeolite and lime soda water softening processes have been improved by the adoption of reaction tanks, settling basins and mech. mixers and conveyers.

C C RUCHHOFF

Philadelphia water supply and health. ANON *Public Works* 61, No 5, 29-30 (1930) —The report of the Philadelphia Hospital and Health Survey Comm. gives the following data concerning the city's water supply. Water intakes are located on the Schuylkill and Delaware rivers. Both rivers receive domestic sewage and trade wastes from numerous municipalities. Av *B. coli* indices at the Schuylkill intakes vary from 10,000 to 34,500 per 100 cc., and at the Delaware intakes the av is 21,000 per cc. Because of a total flood tide travel of 5.95 miles per tide of the Delaware River, the major pollution is caused by the city of Philadelphia. In all there are 29 sewer outlets. The city is operating 2 obsolete and overloaded sewage treatment plants which handle about 10% of the total flow. By-passing of raw sewage into the rivers is common. During the summer months, there was no dissolved O content in either river. The O content of the Delaware intake samples is often 30-40% of satn. Five filtration plants are in operation, but no provision is made for adequate preliminary storage or for prefiltration. All filter plant effluents are chlorinated, the dosage being regulated according to the *B. coli* index of the treated water and not by the residual Cl data. Each plant is operated according to the policy of the individual filter plant superintendent. The water supply has not been approved by the Pennsylvania Department of Health or the U S Public Health Service, since July, 1925. Numerous cases of cross connections were reported. In 1913, 1300 cases of typhoid were attributed to pollution caused by cross connections.

C C RUCHHOFF

Water-supply problems in Holland. F A LIEFRINCK *Public Works* 61, No 9, 19-20 (1930), cf C. A. 24, 3294 —Water has been one of the enemies of the inhabitants of the Netherlands. Much of this water is salt or brackish. Because of unfavorable geo-hydrological conditions difficulties are encountered in providing fresh potable water. Sixty one % of the population, divided over 444 communities, now has water supplies. In rural districts water supplies are built by a combination of several communities. The Government Bureau of Water Supply cooperates and often takes over the financial risks of new enterprises. Of the 444 water supplies 142 are drawn from dunes, 253 take ground water outside of dunes and 49 have river water supplies. The dune supplies are taken from typical lens shaped fresh water accumulations. The water is captured either by means of shallow wells (Haarlem), by open channels (Amsterdam) or by covered drains (the Hague). Future problems arise because of the reclamation of the Zuider Zee which will add some 550,000 acres of fertile land to the country. One of the 4 polders will be ready for occupation in 1931. The soil of these new polders has been under salt water for centuries, so there is little hope of finding fresh water for the population which will settle there.

C C RUCHHOFF

Improvements to the Bluefield water supply. GEO D NORCOM *West Va Univ. Eng Expt Sta., Tech Bull No 3*, 49-56 (1930) —To increase the quantity of the Bluefield Water Supply a slab-type dam was built to impound the waters from several springs. Storage capacity for water now amounts to 172,000,000 gallons. Daily chem

and bacteriologic examination of the water are made. Alga control is practiced to minimize taste and odors. A. L. ELDER.

Water-treatment works of Iola, Kansas. M. F. HARTER. *Water Works and Sewerage* 77, 284 (1931).—Water for this town, obtained from the Neosho River, is at present settled and chlorinated before distribution. The hardness of the water varies from 1.0 to 5.0 parts per million, principally on account of Ca carbonate, which represents 65–75% of the total hardness, and Mg, which is usually lower than 2.0 parts per million. During low flow, the chloride content is as high as 500 to 750 parts per million, because of the effect of oil-field pollution. Most of the year the chloride content is less than 400 parts per million. The new filtration and water softening plant is designed to treat 2 million gallons per day with provision for extension of the plant facilities. Raw water will be delivered by the present low-service pumps to the mixing chamber below the chem. application room. Lime is applied at the entrance of this chamber, and the water has a retention period of about 9 min. and a velocity of 0.5 ft. per second. The water next goes to the reaction tanks for about 30 min. The discharge from the reaction tanks passes to the existing preliminary settling basins for a 3-day retention period. The water is then aerated and passed in a chamber for carbonation. Aeration is obtained by the use of a Sacramento gravel pump, and CO₂ from a natural-gas processing plant in the chem. head house is distributed through a grid system on the floor of the carbonation chamber. A 20-min. retention period is provided in this unit. Following this the water is treated with alum, and it then goes to the filters. There are 3 filter units, each capable of filtering 67,000 gallons per day. Scraper is provided for 27,000 gallons. Cl₂ is applied at the entrance to the last reservoir. C. C. KROSVY.

Present-day water-treatment processes. A. E. CLARK. *Water Works* 61, No. 9, 50 (1933).—C. discusses recent standard practice in clarifying, filtering and chlorinating water, with examples from Tennessee practice. C. C. KROSVY.

Treatment of water and its effect on ferrous incrustations. ARTHUR GOERTZ. *Water and Water Eng.* 33, No. 307, 121 (1931).—The primary cause of Fe oxide incrustations are solids of the water, dissolved Fe, Fe bacteria and dissolved O. Dissolved Fe may be partially precipitated and removed by aeration and filtration, which also tend to eliminate dissolved O where present in excess under pressure. The ultimate source of the incrustation is dissolved Fe in the water itself. Addition of lime or other alkali to water converts solids, but only slightly retards bacterial action. Addition of lime in sufficient quantity to be effective is costly and makes it undesirable for certain technical processes. No useful results are seen from the addition of Na silicate until prohibitive costs are reached. This seems true, in general, of addition of chemicals. Filtration removes only a part of the Fe bacteria. The only sure means of preventing re-deposition after scraping is to treat pipes after scraping. ALICE W. ELLISON.

Water of wells and rivulets of the northern bank of Lake Balaton. G. RUDOLF. *Magyar Kém. Hat. Int. Munkák* 3, 124–127 (1931).—It starts from Tihany to Ácsd were examined. Wells mostly originate from cracks of dolomite or limestone, therefore the salts consist of Ca and Mg HCO₃. The water of the rivulets contains less salt, since bicarbonates are precipitated. Data of 12 detailed analyses are given. S. S. DE FREY.

Floating matter contained in the Danube water at Budapest. T. TAKÁCS. *Hydrokémia* 10, 53–67 (1931).—Water of the left part of the Danube contained 0.2198 g. l. floating matter (av. of 97 detms.) that of the right part only 0.1874 g. l. The av. of the 2 is 0.2036 g. l. The Danube water above Budapest (at 1 point) contained 0.2002 g. l. floating matter. Drains proved to be of no influence on the floating-matter content. The above data mean that the floating matter content of the Danube amounts to about 10,000,000 metric tons a year. Analyses of the floating matter gave: SiO₂ 1.06, TiO₂ 0.21, Al₂O₃ 3.71, Fe₂O₃ 0.54, FeO 1.27, CaO 53.23, MgO 6.45, P₂O₅ 0.24, Na₂O 0.52, K₂O 0.04, MnO 0.04, water (driven off above 110°) 5.74, CO₂ 28.79% (total 93.43%). The large content of CaCO₃ was caused by precip. of bicarbonates during evapn. of the soln. Mech. analyses gave: clay (under 0.002 mm.) 10.23, silt (1.00–0.002 mm.) 53.13, fine sand (0.2–0.002 mm.) 36.32, sand (0.2–0.2 mm.) 13.33%. Microscopical examn. showed that it consists mostly of crystals of calcite. In addition, quartz, biotite, phlogopite, orthoclase, microcline, green and brown fragments of amphibole, ilmenite, garnet, pyrite and hypersthene were found. S. S. DE FREY.

Efficiency of water-purification plants on the Great Lakes. H. W. FREESTER. *Water Works and Sewerage* 77, 285 (1931) and C. A. 25, 139.—The lake system studied provides approx. 10,000,000 people with drinking water and receives sewage and industrial waste equal to a population of five and one-half million persons. There is a striking similarity in all plants in regard to the existence of a well-defined relationship between the bacterial quality of the raw water and the corresponding quality of the effluent at each

successive stage of purification. The marked variability in av. bacterial efficiency in the different plants was found to be the outstanding dissimilarity. In comparing the bacterial efficiency of the Great Lakes plants with Ohio River plants it was found that the former were slightly less efficient with chlorination included and considerably less efficient with chlorination excluded. It is suggested that these divergences might be due to the differences in the chem. compn. of the 2 waters. The study further indicates that a *B. coli* index of 4500 per 100 cc. represents an upper limit of permissible pollution. With av. indexes ranging below 1000 per 100 cc. the majority of plants would be overburdened for a very small proportion of the time. The following areas are listed in their order of decreasing intensity of pollution: extreme southern end of Lake Michigan, extreme western end of Lake Erie, outlet of the Detroit River and southern shore of Lake Erie between Cleveland and Sandusky. C C RUCHHOFF

Laboratory-control tests and their practical significance. PERKINS BOYNTON W Va Univ Eng Expt Sta, *Tech Bull* No 3, 103-6(1930)—Perkins emphasizes the necessity of lab-control data. E.g., the water of the West Fork River coagulates best between a *pn* of 6.2 and 6.8. Detns. of turbidity at 3 hr. intervals make it possible to predict the amt. of alum needed. From 5 to 10 parts per million of alkalinity should remain in the water after treatment. A L ELOER

Analysis of water from the Odiel river (Huelva). L BLAS *Anales soc. españ. fis. quim.* 29, 162-3(1931)—High mortality of fish in the Odiel River led to analyses which showed 0.45 mg. As_2O_3 per l. and *pn* 4.7, quite sufficient to explain the mortality and migration of fish. E M SVENNES

Chemical analysis of the water of "Margit" well of Zánka-Vétkút (Hungary). F MARSCHALL *Hydrol. Közlemény* 10, 136-40(1930)—The dissolved matter totalled 1707.87 mg./l. S S DE FVÁLY

Chemical analysis of the waters of Cave Aggtelek (Hungary). R MAUCHA *Hydrol. Közlemény* 10, 201-7(1930)—The data for analyses of 6 samples are given. The content of org. matter is small. S S DE FVÁLY

Chemical analysis of Szent István mineral water of Pará (Hungary). F MARSCHALL *Hydrol. Közlemény* 9, 100-3(1929). S S DE FVÁLY

Chemical analysis of the temperate springs of Dunaalmás (Hungary). K EKSZT *Hydrol. Közlemény* 9, 104-6(1929)—The waters of 2 springs were analyzed both belong to the group of earthy, simple thermal waters contg. H_2S and HCO_3^- . S S DE F

Attempt to control cyclops in a water plant. E M JOHNSON *J. Am. Water Works Assoc.* 23, 582-5(1931)—Treatment of filters with a concd. soln. of $CuSO_4$ and prechlorination for 1 hr. with 15 parts per million Cl_2 was ineffective. To date an effective dosage or treatment has not been found. D K FRENCH

Determination of the nitrate content of drinking water. Ö SZAKÁCS *Kísérlet. Közlemények* 33, 330-5(1930)—A rapid method consists in adding Ph_3NH in H_2SO_4 to 10 cc. water, shaking the mixt. in a bottle with glass stopper, cooling the liquid quickly and within about 5 min. comparing it with a standard soln. made of basic Cu carbonate, $K_2Cr_2O_7$, and rosolic acid equiv. to 60 mg. N_2O_5 per l. The standard soln. should be diluted to the concn. desired. S S DE FVÁLY

Pre-ammoniation at Springfield, Illinois. CHAS. H. SPAULRING *Public Works* 61, No 5, 35-6(1930).—See C A 25, 1607. C C RUCHHOFF

Combined nitrogen in rain water. C SRIKANTIA *Half-Yearly J. Mysore Univ.* 4, 195-8(1930)—Sixty-two samples of rain water collected in Bangalore over a period of 1 year were analyzed for NH_3 , N_2O_5 , and N_2O_3 . The amts. found show a max. in April. The total rainfall was 32.40 in., giving 1.37 lb. NH_3 per acre, 0.861 lb. N_2O_5 + N_2O_3 per acre and a total of 2.40 lb. of combined N per acre. GEORGE CALINGAERT

The taste problem solved. JOHN R. BAYLIS *Water Works and Sewerage* 77, 299-304(1930)—The elimination of tastes in water has been along 2 lines. One is to prevent the taste from being more pronounced when Cl_2 is added to the water, and the other is to eliminate all objectionable tastes by removing them from the water by changing them to inoffensive compds. The success of the superchlorination treatment depends upon the power of the Cl_2 to break down the offensive compds. and also on the effectiveness of the dechlorinating process. The most common tastes, such as those produced by microorganisms and by the phenolic compds., are readily changed to inoffensive compds. with Cl_2 . With superchlorination all bacterial life is destroyed, the oxidizable org. matter is reduced and the water is more likely to remain stable when stored. SO_2 is a very effective dechlorinating agent. In practice, it is best to add an amt. of SO_2 equivalent to residual Cl_2 present. Na bisulfite is also used for dechlorination. The most successful treatment for the removal of tastes is the Cl_2 activated-carbon process. The water is superchlorinated, and a residual Cl_2 content of 0.5 to 1.0 part per

million, prior to dechlorination, is sufficient if there has been 3-4 hrs. contact time. Charcoal, lignite and activated C are dechlorinating agents. Activated C is best suited for the purpose and removes the excess Cl as well as objectionable compds. not changed by the Cl. Another process is the NH_3 -Cl treatment. NH_3 is added to the water prior to Cl. This is done to prevent chloro tastes. The Cl unites with the NH_3 to form a chloramine, a weak oxidizing agent which does not readily oxidize the org. material present in the water. Chloramine is a more effective bactericide than Cl. It prevents aftergrowth, it eliminates objectionable odors and tastes, it allows a 50% reduction of the amt. of Cl required for actual sterilization. It retards the oxidizing action of the Cl and prevents its absorption, thus allowing practically all the Cl to be available for germicidal action. More time is required to sterilize the water with chloramine than with Cl. KMnO_4 has been used to eliminate taste-producing compds., but it gives the water a slight bitter and astringent taste. Ozone has not been used extensively to eliminate tastes because of its cost, but it may be quite effective. C. C. PIERCE

Substances producing taste in chlorinated water. I. B. A. ADAMS. *Water and Sewerage* 33, No. 367, 116-13 (1931).— CH_3 taste in chlorinated water may arise from the presence in the water of (a) certain products of destructive distn. of coal, (b) certain saliv. derivs. of natural origin, (c) unknown products of bacterial metabolism or decomposition, and of purely chem. action, (d) solids of natural occurrence. Taste may be due to a no. of org. substances, confined to the lower monohydric phenols of the benzene series, their monocarboxylic aldehydes, alcoholic derivs. or their inorg. salts. Tastes more nearly resembling those met with in actual practice are produced by biphenyl monohydric phenols with substituted groups in the ortho position or many solids. Para compds. yield tastes resembling CH_3 , less, but of greater pungency than the tastes of meta compds. The strongest tastes are produced by phenol, *o*-cresol, salicylaldehyde and *o*-hydroxybenzal. The tastable products are probably mixts. of *o*- and *p*-chlorophenol, trichlorophenol or their homologs. As *o*-chlorophenol possesses the strongest taste and more nearly resembles CH_3 , it may be said that the substances producing most intense tastes yield a greater proportion of *o*-chlorophenol.

ALICE WHITTON ELLERSON

Prevention of chlorophenol tastes. CHAS. R. COX. *Public Works* 61, No. 5, 77 (1931). Cf. C. A. 24, 1166.—Expts. concerning the effectiveness of superchlorination, NH_3 treatment and KMnO_4 treatment disclosed the following. Superchlorination is effective but expensive. NH_3 when used before or with Cl prevents formation of chlorophenols. KMnO_4 is feasible and economical. It is necessary to compensate for its influence upon the results of the *o*-tubidine test for residual Cl. C. C. R.

The determination of total alumina in filtered water. F. O. BALDWIN. *Water Works and Sewerage* 77, 311 (1931).—The new method for the detn. of unpyrid. alum is simple, rapid and sensitive to 1 part in ten million as Al_2O_3 . The reagents required for the test are: (1) Three tenths % aq. soln. of alizarin red S (alizarin sodium monosulfonate) in 2 N H_2SO_4 . (2) Saturated soln. of NaHCO_3 . Filter through alumina free filter paper and add 2 cc. dist. water per 100 cc. of filtrate. The strength of this soda soln. should be held close to a normal soln., that is, 5.4% NaHCO_3 (Arm & Hammer bicarbonate may be used). (3) Fifty % AcOH made with the highest grade reagent and boiled dist. water. Use the following procedure. To 50 cc. of the freshly collected and well-shaken water add 1 cc. of No. 1. Use phosphoric acid flasks, size 200 cc. Boil 2 min. on the hot plate, neutralize with No. 2 to just barely purple by adding the soda soln. very carefully drop by drop, then add exactly 0.5 cc. excess, boil 2 min. and cool in cold water. When cold, add 1 cc. of No. 3, let stand 1 min. and compare with the color standard sealed in pyrex 200 cc. flasks similar to the above flask. These color standards are sealed in partial vacuum and have to be standardized very accurately. C. C. PIERCE

Spectrographic determination of cations of some Spanish medicinal mineral waters. I. S. PIÑA DE PÉREZ and C. SIVIENT D'ARMENT. *Anal. Soc. Espan. Fis. Quim.* 29, 255-46 (1931).—Spectrographic tests in the ultra violet region of dry residues from 40 medicinal mineral waters identified Pb, Cu, Ag, Sn, Mo, Ga, Ge and Ti, in addition to the usual elements. E. M. SPENCER

The relation between the fixed residue and electrical conductivity in mineral waters. E. BOVALINI and L. VALASE. *Zur chem. applicat.* 21, 51-75 (1931).—To approximate the compns. of natural waters a large no. of solns. were prepd. of the following 6 types: (1) from $\text{Ca}(\text{HCO}_3)_2$ alone, (2) from SO_4 with NaOH , (3) from SiO_2 , $\text{Ca}(\text{HCO}_3)_2$ and different amts. of NaOH , (4) from SiO_2 , $\text{Ca}(\text{HCO}_3)_2$ and different amts. of NaHCO_3 , (5) from SiO_2 , $\text{Ca}(\text{HCO}_3)_2$ and different amts. of Ca and Mg as sulfate, (6) from SiO_2 and $\text{Ca}(\text{HCO}_3)_2$ plus various salts. The fixed residue at 180° and E at 18° were detd.,

and the ratio $r = \text{fixed res.} / K$ was calculated. Results for r varied from 584 to 4100. The only uniformity is observed for waters of the same type but varying concn., in which case r varies less than 20%.

A. W. CONTIER

Influence of water plants on the electrical conductivity of Balaton water and of bicarbonate solutions. Gy. LUDOVY *Magyar Biol. Kut. Int. Munkái* 3, 482-5 (1930).—Water plants cannot decrease K_{18} ($= 4.5 \times 10^{-4}$) of Balaton water to $0.5-0.8 \times 10^{-4}$. Five g. Potamogeton was used in 100 g. water, and K was measured in 3 hrs. by Plessner's app. The value of K at first decreased, then increased in diluted Balaton water. An increase of cond. was found under the influence of light at far higher K values than those obtained by Ruttner in the water of Lunzer See. The cond. under the influence of light at first increased, then decreased in $\text{Ca}(\text{HCO}_3)_2$ solns. Plants liberate CO_2 in the absence of light. Ruttner's phenomenon was observed on the 3rd day, under pH 8.5 and $K_{18} 0.9 \times 10^{-4}$.

S. S. DE FRIJLY

Protozoological examination of water. Mme N. L. WISAUT *Chem. Weekblad* 27, 526-9 (1930).—Tabulations of the nos. of bacteria (*B. coli*) and protozoa present in previously sterilized waters at successive intervals after seeding show that in most cases both increase at first, the latter much faster than the former, after the period of increase both decrease. The conclusion is drawn that the protozoa can entirely eliminate the bacteria in most cases. Comparative tests of the no. of protozoa and bacteria present in a water may be used to det. the date of infection of the water with bacteria.

B. C. A.

Direct method for the quantitative study of bacteria in water, and some considerations on the causes which produce a zone of oxygen minimum in Lake Glubokoje. S. I. KUZNETSOV AND G. S. KARINKIN *Zentr. Bak. Parasitenk.* 2 Abt. 83, 169-74 (1931).—The diminished content of O in the water below the temp. leap can be easily explained by its continuous consumption by bacteria. The greater d. of these water layers precludes their rising to the surface in consequence of wind circulation. Thus the O in them cannot be renewed by that of the atm. The sharp increase in d. of the water of the temp.-leap layer contributes to the retention of the dead organisms of the plankton. This favors the increased development of bacteria.

JOHN T. MYERS

Gas production and pH determination of coli-aerogenes cultures in sugar broths. C. C. RUCHHOFF, J. G. KALLAS AND BEN CHINN *J. Am. Water Works Assoc.* 23, 565-81 (1931).—Buffering broths with 0.2% K_2HPO_4 is recommended to control pH and increase gas production. The effect of the initial pH values with different media is stated.

D. K. FRENCH

Private cross-connections and similar menaces to the quality of water. JOHN I. CONNOLLY. *Water Works and Sewerage* 77, 309 (1930).—To prevent water pollution through submerged inlets to fixtures and direct waste connections, all water pipes and waste pipe installations must be carefully supervised. Safer fixtures should be developed, and devices to prevent back pressure provided. All direct water connections to waste pipes should be severed. Submerged inlets should be removed, and the water should enter over the rim of the fixture, or a non-siphoning device should be installed.

C. C. RUCHHOFF

The use of sea water for the regeneration of base-exchange materials in water softening. HORACE INGLESON AND BASIL A. ADAMS *J. Soc. Chem. Ind.* 50, 123-4T (1931).—Sea water contg. $\text{Na}^+ 1.092$, $\text{Mg}^{++} 0.127$, $\text{Ca}^{++} 0.046$ and $\text{K}^+ 0.043\%$ was used to regenerate glauconite "Kenzeite." It was possible to substitute sea water for NaCl solns. usually used to regenerate water softeners. Further studies will be made to det. whether or not a film of CaSO_4 forms on the sand during the regeneration process. It was considered that sulfate ions in the sea water might ppt. the Ca ions during treatment of the filters with sea water.

A. L. ELDER

Development of railway water-supply practice. C. R. KNOWLES *J. Am. Water Works Assoc.* 23, 481-94 (1931).—An historical review is given of what has been done, with an outline of the organization and purpose of the American Ry. Eng. Assoc. and the Natl. Boiler Feed Water Comm.

D. K. FRENCH

Boiler water chemistry. W. G. LEBMAN *J. S. African Chem. Inst.* 14, 22-32 (1931).—A review.

E. H.

Boiler feed waters. F. W. W. BUTLER *J. S. African Chem. Inst.* 13, 7-20 (1930).—A review.

E. H.

The need of chemical supervision of boiler feed water. HEINRICH PROSKER. *Chem.-Ztg.* 55, 226-7 (1931).—An analysis of feed water showed the following: residue on evapn. 216.4 mg/l., residue on ignition 152.4 mg/l., loss on ignition 64.0 mg/l., NaCl 2.46 mg/l., SO_2 content 38.76 mg/l., SiO_2 13.2 mg/l., CaO 54.4 mg/l., MgO 17.09 mg/l., total hardness 7.38, carbonate hardness 3.50. The water was clear, free from Fe and Al_2O_3 , gave a neutral reaction with litmus and a weakly alk. one with phenol.

phthalein. The preheater tube of a boiler supplied with this water became clogged because of a hard deposit, 95.61% of which was sol in hot water. The water soln of this scale presented the following compn: Na_2SO_4 67.45, Na_2CO_3 29.18, NaHCO_3 0.14, $\text{Ca}(\text{SO}_4)$ 1.20, NaCl 1.02, SiO_2 0.30, CaSO_4 0.34, MgSO_4 0.00, H_2O 0.31%. R. C. F.

Modern boiler feed water treatment methods. A. I. WARNER. *Chemical Markets* 28, 279-83 (1931).—The use of Na aluminate has improved water treating methods. Diagrams of exptl glass boilers are shown. A. L. FIDLER.

Boiler feed water treatment in Great Britain. A. W. CHAPMAN. *J. Am. Water Works Assoc.* 23, 547-50 (1931).—A review involving softening, embrittlement, naturally soft surface waters and high pressure operation. D. K. FRENCH.

Embrittlement and protection of steam boilers. FREDRICK G. STRAUB. *Power Plant Engineering* 35, 100-3 (1931).—See C. A. 25, 1311. G. L. KELSO.

Prevention of boiler scale. G. V. BÖLSCHÄTZ. *Technika* 10, 325-6 (1929).—Ca is added to the boiler in the form of coned lime water. A 10% soln of Na carbonate is added separately. The deposit should be removed daily. Alk. water destroys varnish insulations of heating tubes, therefore rubber or soft Cu rings should be used for insulation. S. S. DE LINALY.

Blow-down losses and the means for their correction. A. R. MOURAO. *Pulp Paper Mag. Can.* 31, 495-7 (1931).—A discussion is given of the reasons for the losses which occur in boiler blow-downs, and the means for their correction. From a discussion of the effects of surface tension and concn of sol salts and insol matter, it is concluded that the function of the blow-down valve is the reduction of concn of sol salts, and that the reduction of concn of suspended matter is incidental and is efficient only to a degree. It is further concluded that reduction in blow-down is dependent upon the efficient elimination of suspended matter and the reduction of sol salts of the feed water. Conditions which existed in several actual cases and the steps taken for the correction of the losses involved are described. A. PAPINEAU COUTURE.

Unusual corrosion problems. F. B. PORTER. *J. Am. Water Works Assoc.* 23, 534-7 (1931).—Two cases of corrosion, one attributed to electrolysis and one to disid. water action at high temps. are described. D. K. FRENCH.

Sewage treatment at Schenectady. MORRIS M. COUV. *Public Works* 61, No. 3, 30 (1930).—A description of the plant operation for 1929 is given. Because of the inadequate capacity of the plant, anal. of 0.61 m. g. d. of sewage was wasted directly to the river. Trash collected from the bus station is plowed under. Sludge is dried and sold to farmers at 25 cents per load. The demand has exceeded the supply. An analysis of dried sludge shows 1.01% C, 1.22% N and 6.4% ether insol matter. A manual-controlled soln feed chlorinator is used for the prevention of odors during the summer months. The av. prechl. disinfection dosage was 4 p.p.m. Sludge analysis indicated the following averages: 44% moisture, 65.9% org. solids and 68% fly life on the sprinkling filters was controlled by flooding for a period of 21 hrs. C. C. RUCIMORR.

West Side sewage-treatment works, Sanitary district, Chicago. *J. Western Soc. Eng.* 36, 40-110 (1931). General problem. LANGDON FRANK. *Ibid.* 90-2. Chemical and biological features. I. W. MOHLMAN. *Ibid.* 92-4.—Chem. settling makes it possible to decrease the amt of air required in activated sludge treatment and results in a substantial removal of the heavier sewage solids with a decrease in over-all vol. of sludge. Design of the West Side sewage-treatment works. L. C. WHITTEMORE. *Ibid.* 94-7. Mechanical engineering features. HOMER I. STEFFA. *Ibid.* 97-100. Construction methods. L. B. BARKER. *Ibid.* 100-5. Electrical engineering features. JOHN T. HAWLEY. *Ibid.* 105-7. Operation. A. H. GOODMAN. *Ibid.* 108-10.—The beginning of the world's largest sewage purification plant has proceeded with only minor operating difficulties. W. H. BOYNTON.

Pasadena's sewage-disposal plant. G. K. HOORNS. *Municipal Sanitation* 2, 174-6 (1931).—This plant serves a combined population of 121,000. The activated-sludge process is used, preceded by an equalization tank and grit chamber. Part of the effluent is diverted for irrigation purposes, and the remainder discharged into a stream. The sludge is converted into a com. fertilizer. Fifty lb of this material per annum per unit of population served is obtained. Receipts from sales of fertilizer practically cover labor costs of operating the entire plant. G. L. K.

Barrington sewage-treatment plant. E. BENDER. *Public Works* 61, No. 9, 33 and No. 10, 21 (1930).—The sewage of Barrington, N. J., is clarified and treated by mech. aeration. The settled effluent is filtered through sand and chlorinated. The sludge-digestion tanks and covered sludge beds are provided for digesting the activated sludge. C. C. RUCIMORR.

Sewage chlorination for the protection of masonry sewers against disintegration.

LINN H. ENSLOW *Water Works and Sewerage* 77, 306(1930) — Sewers are often attacked by the H_2SO_4 produced in septic sewage by bacteria. This can be prevented by chlorinating the sewage, the chlorination suppressing biological action and simultaneously preventing odors and reducing the O demand. Cl is applied during max. sewage flow, about 9 to 12 hrs. daily. The quantity of Cl used is about 5 parts per million.

C C R

Laboratory studies of sludge digestion. A. M. BUSWELL AND S. L. NEAVE. Ill. State Water Survey, *Bull.* No. 30, 44(1930) — The major nitrogenous constituents were shown to be urea, nucleoproteins and simple proteins, and their degradation products. The hydrolysis of urea to NH_4 , carbonate was rapid and complete. Insol. proteins were hydrolyzed and peptized slowly in acid sludges (pH 5.0) but rapidly above pH 6.4. Fatty acids, NH_4 and free amino acids were the normal products, but amines also were produced in acid sludges. Fatty acids were decomposed, with the production of CH_4 and CO_2 . Amino acids were in part resynthesized into bacterial protein. The purine fraction of the nucleoproteins suffered almost complete destruction. During the remobilization in forms of N, not more than 5% of that present escaped as gaseous N. NH_4 , living bacteria and amino acids were the main nitrogenous end products. Besides proteins, the sewage solids were shown to contain grease (soaps and fats), crude plant fiber and humus-like bodies. Grease degradation was demonstrated, and lower fatty acids, CH_4 and CO_2 were found as end products, thus grease was proved to be an important source of CH_4 in digestion tank gases. The degradation of lower fatty acids was studied to determine the biochemical mechanism involved. For the degradation of lower fatty acids (and by inference for higher ones) an anaerobic oxidative mechanism has been demonstrated in which water acts as the oxidizing agent. Methane and CO_2 were produced in nearly theoretical yields. This mechanism probably has the same wide application in anaerobic processes that the β oxidation mechanism has in aerobic metabolism.

F W TANNER

Chemical studies on sludge digestion. S. L. NEAVE AND A. M. BUSWELL. Ill. State Water Survey Division *Circ.* No. 8, 3-9(1930) — See preceding abstract F W T.

Two questions on sludge digestion answered. JOHN R. DOWNES *Water Works and Sewerage* 77, 217-8(1930) — D discusses the necessity of using circulating pumps for heating coils of sludge-digesting tanks and also the necessary pitch for the floor of digestion tanks. Pumps are preferable to secure equiv. heat transfer at a lower initial water temp. The lower water temp. is necessary to prevent baking sludge on the outside of the heating coil. D recommends a 60° slope for the bottom of the digestion tanks when sand and grit are present. If grit chambers and screens are used, a slope of 18 in. in 19 ft. is sufficient.

C C RUCHMORT

Molds utilized to make sewage sludge inodorous. CHESTER G. WIGLEY. *Public Works* 61, No. 5, 38(1930) — Sewage sludge from sedimentation tanks having a detention period of 45 min. is pumped onto sludge drying beds. When dried so as to be shoveled easily, it is removed and spread for further air drying. It is then ground or broken up and stored in a dark building where mold growths complete the deodorization. The ground material is used as a fertilizer and sells for \$20 per ton. The analysis after treatment is N 1.96%, available N 0.76%, P_2O_5 0.62%, potash 0.13% and organic matter 35.9 per cent. The process has been patented.

C C RUCHMORT

Trickling filters. JOHN R. DOWNES *Water Works and Sewerage* 77, 313(1930). — Different factors must be considered. The type of influent, the depth of the filter and other factors det. to a large extent, the operator's method. In general, the sewage should come to the filters as fresh as possible and with the least possible suspended matter, which will provide against filter-clogging. This keeps down odors at the sprays and passes on to the sprinklers the fine particles which can be better treated aerobically in the filters than anaerobically with the sludge. Clogging of trickling filters is more likely to be due to growths on the bed than to the sewage solids themselves. Such growths can be eliminated by chlorinating or using bleaching powder. Odors also can be controlled by the use of Cl in sufficient quantities to oxidize the H_2S . The other objectionable feature of trickling filters is the presence of the filter fly. The only method of control at present is that of submerging the filter media at regular intervals, by which means the eggs, larvae and pupae are washed out of the filter bed. The adult flies are best destroyed by spraying the walls and walks with kerosene.

C C RUCHMORT

Present practice in industrial-waste treatment. ERNEST W. STEEL. *Public Works* 61, No. 9, 21(1930) — A review is given of the methods used for treating milk wastes, beet sugar waste, corn product wastes, cannery wastes, packing house wastes, tannery waste, textile and wool washing wastes, paper mill wastes, coke by product wastes, acid wastes and mine wastes. The methods suggested for various wastes include

coagulation, sedimentation, treatment on sprinkling filters or coke filters and activated sludge treatment. It is pointed out that there is no satisfactory treatment for wool-scouring wastes. C. C. RICHMONT

Treatment of trade wastes, a necessary feature of stream-pollution control. W. W. HOBBS. W. Va. Univ. Eng. Expt. Sta., *Tech. Bull.* No. 3, 73-83 (1930).—A review is given of some of the research projects which have led to the recovery of valuable products for industry and at the same time decreased the pollution of streams into which industries have been dumping wastes. A. L. ELDER

An innovation in meat-packing waste treatment. H. O. HALVORSON. *Municipal Sanitation* 2, 166-70 (1931).—Biol. treatment of sewage from a meat packing plant killing 5000 hogs daily was found to be too costly and did not allow for protein recovery. The waste from this plant carried \$120-300 worth of N per day. Chem. pptn. methods were unsatisfactory because of the high cost of filtering and drying the ppt. By use of CaSO_4 of the org. N was pptd., and suspended solids were reduced 95%. The best results were obtained with the strongest sewage. The sewage is subjected to a short period of settling, a thorough mixing with Ca and final sedimentation in a Dorr thickener. From 1500 to 1700 lb. of Ca are used per day. Five tons of sludge are recovered per day from 750,000 gallons of sewage. The dry sludge will contain 8.5% NH_4 . At \$30.00 per ton for the sludge the plant will be self-supporting. G. L. KILSO

Stream-pollution investigations in West Virginia. L. KERMIT HERNDON. W. Va. Univ. Eng. Expt. Sta., *Tech. Bull.* No. 3, 68-74 (1930).—The activities of the State Water Commission are outlined. A report of the pollution of Cheat River indicated the pollution of the river from the following sources: 60 coal mines discharging acid waters, a tannery with a biochem. O demand equiv. to 13,000 people, 7 sawmills, a wool washing plant, cement factory, blanket factory and domestic sewage from 10,000 people. No satisfactory program for eliminating the pollution from coal mines is available at this time. A. L. ELDER

Atmospheric pollution in 1929-30. JOHN B. C. KERSTRAW. *Engineer* 151, 257-9 (1931).—A review of reports on conditions at Hamburg, Germany, and at Southampton and Salford, England. D. B. DILL

Hygienic importance of sulfur, of sulfuretted waters and sulfur baths. MAX GATNEWALD. *Pharm. Presse, Wiss. Fachl. Hef.* 1931, 38-40.—A general discussion is given of the element in its varied forms and combinations, its therapeutic application and reputed effect on the animal economy. W. O. E.

Solving the garbage problem in Alton, Ill. THOMAS BUTLER. *Public Works* 61, No. 5, 51 (1930).—By the adoption of a municipal system of rubbish and garbage collection and disposal, a savings of \$35,000 per year was effected. C. C. RICHMONT

Process water in the pulp and paper industry (SPILTINGER) 23. Treatment and disposal of wool washing effluent (SMITH) 25. Treatment of suint liquors from wool-scouring (KING) 25. Coating concrete [reservoirs] (Fr. pat. 698,562) 20. Fertilizer from distillery and sugar factory residues (U. S. pat. 1,799,176) 15. Recovering organic substances from aqueous solutions [in purification of water] (Ger. pat. 522,480) 13.

NURSE, C. J. Purification and Disposal of Sewage. London: Crosby Lockwood and Son, 4s. 6d. Reviewed in *Water and Water Eng.* 33, 187 (1931)

Composition for preventing boiler incrustation. HERMANN JURESCH. Austrian 121,756, June 15, 1929. The compn. comprises powd. gallnut 5, powd. pine bark 2, powd. larch bark 2 and resin or grafting wax 1 part. 100 g. of the compn. are used for 1 cubic m. of very hard water.

Sewage-disposal apparatus. WM. W. SAYERS and MARCUS B. TARK (to Link Belt Co.). U. S. 1,799,150, April 7. Structural features.

Filter suitable for filtering water as drawn for use. ANTON C. MENGE. U. S. 1,800,093, April 7.

Sewage treatment. FREDERICK W. SPERR, JR. (to Koppers Co.). U. S. 1,799,444, April 7. Sewage is aerated in the presence of a compd. of a metal, the sulfide of which is insol., such as $\text{Fe}(\text{OH})_3$, and a foam is maintained on the surface of the liquid during the aeration. App. is described. Pine oil may be used as a flotation agent.

15—SOILS, FERTILIZERS AND AGRICULTURAL POISONS

J. J. SKINNER AND M. S. ANDERSON

Agriculture and soils of the coastal plain of Palestine. Soil survey of the Jaffa Sub-district. H. E. Z. RACZKOWSKI *Palestine Dept Agr and Forests, Agr. Leaflets 11 Ser., Soil Survey No. 1*, 24 pp (1929). *Expt Sta Record* 63, 17-8.—The usual soil survey information is given, with the addition of some detailed chemical and mechanical analyses given in an appendix. G. G.

Investigation of soils and of the mineral content of pasture grasses at Waranama Ranch, Berbice River. R. R. FOLLETT-SMITH *Agr J Brit Guiana* 3, 142-59 (1930).—Both soils and pastures examined were deficient in CaO, phosphate and potash, and grazing cattle became quickly emaciated. The feeding of bone meal to cattle is recommended. B. C. A.

Negative values in Neubauer-Schneider tests of soils. M. GRAČANIN *Z Pflanzenernähr Düngung u Bodenk* 18A, 115-8 (1930).—Negative values for phosphate assimilation in Neubauer tests are ascribed to the transference of phosphatic material from the germinating seeds to the soil and its subsequent fixation by the soil. The phenomenon is characteristic of highly adsorptive soils deficient in phosphate. B. C. A.

Water condition of the soil and plant growth. HANS GRADMANN *Naturwissenschaften* 19, 237-64 (1931).—A review dealing with the mechanism of utilization of soil water by plants, mobility of water in soil, etc. Many references are given. B. J. C. VAN DER HOEVEN.

Soil moisture content at which barley wilts. L. SMOLIK *Bull Czechoslov Acad Agr* 1930, 166.—Barley plants can utilize, at least partly, the hygroscopic moisture of soils, particularly in heavy soils. B. C. A.

The problem of the accuracy of soil moisture determinations. D. S. KUZNETZOV *Zhur Opit Agron Yugo Vostoka (J Expt Landw Südost Eur Russlands)* 6, No 1, 155-80 (1928), *Expt Sta Record* 63, 716.—K. presents the results of a mathematical treatment of data on the moisture of a soil under sunflowers. The moisture content was determined on each 10-cm layer to a depth of 200 cm. For each layer 8 borings were made, so that each plot gave 160 determinations. The genetic horizons in the soil investigated were found not to influence the moisture content. From a mathematical point of view it was more rational to determine the moisture on each 10-cm layer. A sharp break in the moisture content and its constancy could be noted at a depth of 100 cm. It is important to note the presence of a correlation between corresponding values of moisture in 2 different 10-cm layers on the one hand and 2 different borings on the other. The correlation value of the moisture in two borings is determined by the distance between them. The character of the change in the accuracy of individual measurements of moisture within the m layers (0-100, 100-200 cm) or of independent experiments is analogous to the character of the change in the constancy of the moisture, both the accuracy and constancy being measured by the same mean quadratic departure. G. G.

The coagulation of aqueous suspensions of soil with barium sulfide and calcium sulfide. B. G. ZAPROMETOV *Pochrovedenie* 24, No 3-4, 37-42 (in German 43-60) (1929).—Z. compared the coagulation powers of CaS and BaS for soil suspensions and for mastic with those of BaCl₂, CaSO₄, CaCl₂ and Al₂(SO₄)₃. If 0.05 N (0.02 N of CaSO₄) solutions were used, a comparative value could be obtained for each of the cations. BaS and CaS were effective as coagulants and were adsorbed by the mastic, BaCl₂ and CaCl₂ were also effective as coagulants but were not adsorbed by the mastic. J. S. JOFFE.

Replaceable bases and water sorption of soils. L. SMOLIK *Bull. Czechoslov Acad Agr* 1930, 10.—The moisture absorbing capacity of soils saturated with different bases varied with the base used in the order Na > Ca > Mg > K > NH₄. B. C. A.

Degree of saturation of the adsorptive complex in Moravian soil types. B. MALÁČ. *Bull Czechoslov Acad Agr*, 1929, 853. B. C. A.

Role of humus in the absorption complex. L. KOTZMANN *Mezőgazdasági Kutatások* 2, 537-55 (1929).—No connection was found between the total amount of organic matter and that of alkali- or NH₄-sol matter determined by the method of Grandjean Hilgard. The NH₄-absorbing capacity of humus preparations made from different soils is the same, indicating an equivalent weight of 30S. The NH₄-absorbing value for soils is much larger than the saturation values determined by the method of Hussink. The conductometric titration value stands between above data. S. S. DE FINÁLY.

Optical method for the investigation of humus. G. I. POKROVSKII *Pochrovedenie* 24, No 1-2, 124-30 (in English 131-6) (1929).—P. presents a photometric method for

the detn. of humus. The method is based on the well known phenomenon of reflection and absorption of light by variously colored, variously dispersed substances or by a difference in compn of the substances examd. A description of the app as well as a math. discussion of the formula employed is given, and data presented check with the chem. methods. J. S. JOFFE

Chemical changes in the organic matter [of soils] during the natural decomposition of the humus layers of woodlands. I. Variation in pentosan content. A. NEMEC. *Z. Pflanzenernähr. Düngung u. Bodenk.* 18A, 65-104(1930)—The decompn of pine, fir and larch needles does not commence until after that of the cell wall substance. In general the pentosan content (ash free basis) decreases as humification proceeds although this relationship is less marked in soils under deciduous trees. Humification precedes pentosan decompn. Matter sol in $\text{CaH}_2\text{-alc}$ is decomposed at a relatively greater rate than is the total org. matter and is greater in pine than in fir soils in a similar stage of humification. The high proportion of resin in the needles tends to inhibit the normal humification process. The consumption of $\text{CaH}_2\text{-alc}$ exts varies directly with the amt. of material extd. The nitrifying power of these soils is closely allied to the amt. of $\text{CaH}_2\text{-alc}$ -sol. matter present, the latter becoming toxic when present in quantities exceeding 5% of the ash free dry matter. B. C. A.

Occurrence and distribution of salinity in a virgin mallee soil. J. E. THOMAS. *J. Council Sci. Ind. Research* 4, No. 1, 12-9(1931)—Water sol. salts and chlorides were detd. at 1 ft. intervals to a depth of 16 ft. At the surface the Cl content in parts per 100,000 was 0.029, at 4 ft. 0.134, at 10 ft. 0.182, and at 16 ft. 0.278, indicating that leaching action had come to the greater depths. The botanical flora of the soil was decidedly halophilic. Diagrammatic soil profiles and mech. analysis are given. The study was conducted with a view to possible future irrigation. C. R. FILLERS

Can a water extract serve as an index to characterize the degree of salt treatment of soils? D. I. TARASOV. *Pochvovedenie* 24, No. 3-4, 61-69 (in English 99-101)(1929)—Analyses on aq. exts. of 5 different alkali soils are reported. They show that H_2O exts. all of the sol. salts, which indicates that the chlorides and sulfates are not in an adsorbed state. The methods used are described in detail. An English summary is given. J. S. JOFFE

The biodynamics of alkali soils. M. I. PRIKHOODO AND M. I. BELIKOVA. *Pochvovedenie* 24, No. 3-4, 145-66 (in English 167)(1929)—On soils in the open field NaCl hindered the activity of *Acetobacter*, whereas Na_2SO_4 stimulated their multiplication, and the amt. of N fixed was just as great as that in the check plots. *Clostridium pasteurianum* developed better on the NaCl soil, especially in the upper horizons. Anaerobic as well as aerobic decompn. of cellulose was also stimulated by treating the soil with salt. Nitrification was stimulated in the soils treated with Na_2SO_4 but not in those treated with NaCl. The alky. of the soils also increased because of the salt treatment and subsequent leaching by rain water. J. S. JOFFE

Repeated soil extractions according to the method of Th. Sadel for the estimation of the solubility of the potassium in the soil. H. LINSOGAARD. *Landes Vers.-Sta.* 111, 261-9(1931)—The Sadel method for the detn. of the soly. of the K in the soil is not of much value for practical use but gives some information as to the mechanism of soln. JOHN R. HILL

Replaceable sodium in soils. K. K. GEDSOIL. *Pochvovedenie* 24, No. 1-2, 1-14 (in English 14-21)(1929)—Ten g. of soil (if it contains more than 0.1% replaceable Na) or 100 g. (if less than 0.1% of Na is present) is mixed with 10% c.p. CaCO_3 . 1000 cc. of H_2O is added and a current of CO_2 is passed into the mixt. for 2-3 hrs. The Na is detd. on the filtrate. If sol. Na salts are present, they are washed out before the soil is treated with the CaCO_3 and CO_2 . The soil is washed until dark streams of humus begin to appear. J. S. JOFFE

The method of determining the lime requirement of the soil. V. M. FILIPENYA. *Udobrenie i Urozhai (Fertilizers and Crops)* 2, 642-5(1930)—By a method similar to that of Yarusov, F. calculates the lime requirement from a detn. of the hydrolytic acidity by the CH_3COONa method. He uses the coeff. 2.0 instead of 1.75, figuring that the acid extd. represents only 50%. F. used the Trenel potentiometer for the detn. of the pH. J. S. JOFFE

A simplified method for determining the lime requirement by the hydrolytic acidity of the soil. S. S. YARUSOV. *Udobrenie i Urozhai (Fertilizers and Crops)* 2, 638-42 (1930)—With the equation (1) $[\text{H}^+] = K[\text{AcOH}]/[\text{AcONa}]$ as a basis, Y. deduces a series of equations from the reactions which take place when a soil is treated with AcONa for the detn. of hydrolytic acidity. From the pH or the $[\text{H}^+]$ it is easy to det. the AcOH concn. in the soln. and, hence, the hydrolytic acidity without titrating the

soln. By substituting the numerical values of formula (1) in terms of mg Ca the following equation (2) is obtained $[H^+] = 1.8 \times 10^{-8} (X/5000)$ or $[H^+] = 3.6 \times 10^{-8} X$ (1.8×10^{-8} being the dissociation const. of $AcOH$), where X is the content of $AcOH$ in 250 cc expressed in terms of mg Ca. Hence (3) $\log [H^+] = \log 3.6 \times 10^{-8} + \log X$ or (4) $pH = 8.4177 - \log X$. In this equation X consists of 2 components: the content of $AcOH$ in the original soln of $AcONa$ before the detn and the $AcOH$ obtained by the reaction of the acid soil with the $AcONa$, or $X = X_0 + X_1$. Substituting for pH in (4) its value 8.2 (the original pH of the $AcONa$) and solving, X_0 (calcd in terms of Ca in 250 cc) = 1.75 mg. H_2A_1 is multiplied by the coeff 1.75 (the Kapper coeff. for podzol soils) the hydrolytic acidity is obtained. In practice the detn is made as follows: A sample of soil (100 g) is mixed with a 1.0 N soln of $AcONa$ (pH 8.2) in the ratio of 1:2.5 and shaken for 1 hr. The soln is then filtered and pH detd. Let the pH be 6.8, then from equation (4) $X = 41$. Subtracting 1.75 mg (X_0) gives $X_1 = 42.25$ mg of Ca. By multiplying this by 1.75 (the Kapper coeff.) the hydrolytic acidity is detd. directly in terms of Ca, and thus the lime requirement is detd. directly without titrating. J. S. JOFFE

The influence of calcium sulfide on the percolation of water through the soil. A. N. RUZANOV. *Pochroedenie* 24, No. 3-4, 41-50 (in German 37-6) (1929) — Expts with CaS as a coagulating agent on heavy clay soils have shown that when applied as the salt it is more effective than $CaSO_4 \cdot 2H_2O$ as measured by the percolation of water. Applications of CaS also decrease the alk. of the soil and increase the SO_4 content and CaO . Adsorption expts. with mystie have shown that the Ca from CaS is adsorbed in a quantity of 80-83%, whereas only 37-39% is adsorbed from $CaCl_2$. J. S. J.

Biochemical investigations on arable soil. A. J. J. VAN DERVELDE, A. VRAEBELIN and L. DEKOKER. *Compt. rend.* 192, 766-7 (1931) *C. r. l. A.* 24, 488 — Three soil factors are detd.: the volumetric acid value, the volumetric alk. value and the volumetric value of adsorption. The volumetric acid value is obtained by shaking 5 g. of soil every hr. for 9 hrs. with 40 cc. of distd. H_2O and 10 cc. of 0.1 N H_2SO_4 or HCl . The flask is then allowed to stand for 15 hrs., after which 25 cc. of the liquid is decanted and titrated with 0.1 N $NaOH$. From this is calcd. the acid retained by the soil or the volumetric acid value. The alk. value is obtained in a similar way by treating the soil with $NaOH$ and titrating with H_2SO_4 . The difference between the amt. of 0.1 N HCl/H_2O_2 retained and 0.1 N HCl or H_2SO_4 retained is calcd. the volumetric value of adsorption. The volumetric acid values and volumetric values of adsorption follow the order obtained in using dyes or centrifuged milk, and these values increase as the concn. of the H^+ diminishes. The volumetric alk. value generally decreases as the corresponding acid value increases. J. R. ADAMS

The direct method of counting bacteria in relation to the adsorption capacity of the soil adsorbing complex. N. N. MALCHYVSKAYA. *Pochroedenie* 24, No. 3-4, 131-43 (in German 143-4) (1929) — By the Vinogradskii method of direct microscopic count, bacteria counts were made on several soils, samples were taken from the resp. horizons. Two counts were made on each soil sample: (1) as the soil appeared in nature and (2) after the soil had been satd. with Na by the Gedroiz method. It was found that by dispersing the soil—through satn. with Na—the no. of bacteria increases. J. S. JOFFE

Colloidal content and related sod factors as indicators of (forest) site quality. I. T. HAIG. *Yale Univ. School Forestry Bull.* 24, 73 pp. (1929), *Expt. Sta. Record* 63, 420-1 — The main purpose of the work was to measure the value of colloidal content as an indicator of site quality, and to make such incidental observation of the value of other soil measures as the character of the data would permit. The conclusions are considered sp. for the brown, slightly podsolized forest soils of southern Connecticut and of certain adjacent territory, and to be applicable particularly to such of these soils as occur commonly in the vicinity of New Haven. Colloidal content and silt-plus-clay content are fair measures for site quality. The org.-matter content varied between 2 and 10% and was without significant influence on soil fertility. The value of soil type as an indicator of site quality cannot be accurately estd. with the data available. A knowledge of the soil class (the textural quality of the A horizon or surface soil) permits classification of soil quality within approx. one broad site class. G. G.

Additional data on the mutual coagulation of colloids in the soil. V. N. SIMAKOV. *Pochroedenie* 24, No. 1-2, 22-77 (in German 78-123) (1929) — S. tested out the crit. concns. of MnO_2 with a const. quantity of $Fe(OH)_3$ at which mutual coagulation would take place. He found that the zone of coagulation is at a point between 0.577 and 1.227 moles of MnO_2 for 1 mole of $Fe(OH)_3$. Mutual coagulation of $Fe(OH)_3$ and SiO_2 takes place within narrow limits: from 2.757 to 3.452 moles of SiO_2 for 1 mole

of $\text{Fe}(\text{OH})_3$. With an excess of $\text{Fe}(\text{OH})_3$ the mixt. becomes pos., whereas with an excess of SiO_2 the mixt. becomes neg. Mutual coagulation of $\text{Fe}(\text{OH})_3$ and Na-satd. soil suspension takes place within a definite zone, for 1 Fe_2O_3 (by weight) it takes 3.753-8.043 unit weights of Na satd. soil. Two zones were noted at which a stable mixt. of the 2 colloids persists either an excess of $\text{Fe}(\text{OH})_3$ or of the soil suspension. These mixts. behave differently towards electrolytes, with an excess of the soil suspension the point of coagulation with an electrolyte is equal to that of a pure suspension. With an excess of $\text{Fe}(\text{OH})_3$ the mixt. is sensitive to the anions. Mutual coagulation of $\text{Al}(\text{OH})_3$ and MnO_2 takes place as follows: It takes 1 mole of Al_2O_3 to 0.085-0.447 of a mole of MnO_2 , or for 1 mole of MnO_2 it takes 2.237-11.765 moles of Al_2O_3 . A comparison with the Fe_2O_3 sol shows that 1 mole MnO_2 will take 0.815-1.733 moles of Fe_2O_3 against 2.237-11.754 moles of Al_2O_3 . Thus, the MnO_2 is 6.8 times as effective for the Al_2O_3 as for the Fe_2O_3 . With SiO_2 the relation is 1 SiO_2 to 0.577-1.302 Al_2O_3 , whereas with Fe_2O_3 the relation is 1 SiO_2 to 0.290-0.293 Fe_2O_3 . With the soil suspension the relation is: 1 unit weight of suspension to 0.238-0.810 unit weights of Al_2O_3 . Just as with $\text{Fe}(\text{OH})_3$, an excess of $\text{Al}(\text{OH})_3$ in the mixts. is sensitive to anions, but the proportions are not the same as for the $\text{Fe}(\text{OH})_3$. From the data obtained calcns. were made for the coagulation values of each sol in terms of another or a combination of them. Thus from the $\text{Fe}(\text{OH})_3$ data 1 mole of MnO_2 = 2.813 SiO_2 , or 1 SiO_2 = 0.355 MnO_2 . From the $\text{Al}(\text{OH})_3$ data 1 MnO_2 = 3.879 SiO_2 , or 1 SiO_2 = 0.258 MnO_2 . By combining 2, 3, 4 or more sols of the same charge, the amt. of the sol of the opposite charge can be calcd. by taking the figures for each individual sol. Numerous examples and calcns. are given.

J. S. JORRE

Proportions of easily and difficultly mobilizable acid within the zone of exchange acidity in soils, and the bearing of buffer values on this. S. GOV, P. MCLELLAN AND O. ROOS. *Z. Pflanzenernähr. Düngung u. Bodenk.* 18A, 104-14 (1930), cf. C. A. 24, 5534.—Differences in the exchange acidity values of soils as detd. by Daikuhara's method (C. A. 9, 550) and by electrometric titration are attributed to the varying proportions of easily and difficultly sol. acids present. The latter are not necessarily recorded by Daikuhara's process. Electrometric titration curves are adapted to show base fixing areas and areas characteristic of easily and difficultly sol. acids. In humus soils there is a greater proportion of difficultly sol. acids concerned in producing exchange acidity than is the case with mineral soils. Buffer values of soil are calcd. from electrometric titration curves (in N KCl) over the range pH 6.8-7.2. Detn. of exchange acidity in soils by this means is limited to mineral soils with less than a prescribed buffer value.

B. C. A.

The colorimetric determination of phosphoric acid in hydrochloric acid and citric acid extracts of soils. R. G. WARREN AND A. J. FUGLE. *J. Agr. Sci.* 20, 532-40 (1930).—The colorimetric detn. of P_2O_5 in HCl and citric acid exts. of soils, by a method involving the evapn. of the ext., ignition and acid extn. of the residue with either the Denigès or Fiske Subbarrow methods of color development, was satisfactory only with light soils. Clay soils gave low results because of the presence of larger amts. of Fe. A method is given in which the org. matter and Fe are removed by treatment with NaMnO_4 and $\text{K}_4\text{Fe}(\text{CN})_6$. For HCl exts. 20 g. of powd. soil (1 mm.) is boiled for a short time with 70 cc. of concd. HCl and allowed to digest on a water bath for 48 hrs. The soln. is then cooled, diluted, filtered and made up to 250 cc. Fifteen cc. of this ext. is treated with 0.5 cc. of 20% NaMnO_4 and kept on a hot sand bath for about 15 min. The liquid (which should now be free of any brown Mn ppt.) is cooled, diluted to about 30 cc., and treated with 6 cc. of 10% $\text{K}_4\text{Fe}(\text{CN})_6$, followed by 5 cc. of 10% MnSO_4 soln., the mixt. is shaken frequently. After standing several min. the mixt. is titrated with NH_4OH until the blue color just turns purple. 3.5 cc. of 2 N H_2SO_4 is added, the whole is transferred to a 100 cc. volumetric flask, made up to vol. and filtered. Aliquots of the filtrate are taken for color development by the Fiske Subbarrow or Denigès method. For citric acid exts. 25 g. of soil (2 mm.) is shaken for 24 hrs. in a half liter bottle with 250 cc. of 1% citric acid (extra citric acid equiv. to the CaCO_3 present being added). The soln. is then filtered. Seventy-five cc. of this filtrate is treated in a 250 cc. Kjeldahl flask with 10 cc. of concd. HCl and 12 cc. of 20% NaMnO_4 , this is allowed to stand $\frac{1}{2}$ hr. and then vigorously digested until no Mn ppt. remains. The mixt. is then transferred with a min. of water to a 100-cc. volumetric flask, and 4 cc. of 10% $\text{K}_4\text{Fe}(\text{CN})_6$ is added drop by drop with shaking. After several min. the mixt. is titrated with NH_4OH as in the method for HCl exts. 1.5 cc. of 2 N H_2SO_4 is then added and the voln. made up to vol. After filtering, the color is developed in an aliquot as with HCl exts. The results are in good agreement with the gravimetric method.

P. R. DAWSON

The phosphoric acid and lime content of some arctic and Norwegian soils. E. G. DOERELL. *Superphosphate* 4, 105-9(1931).—Uncultivated soils from 5 localities in Spitzbergen and Norway contained 0.2-3.8 mg of readily sol P_2O_5 per 100 g. of soil, as detd by the Dirks method. Two samples of soil from potato fields contained 3.8 and 17.4 mg of readily sol P_2O_5 per 100 g. There was no connection between the P_2O_5 content of the soil and the type of rock from which the soil was derived.

K. D. JACOB

Distribution of assimilable phosphoric acid in the soil under influence of factors of production. L. DWORAK. *Kiserlet Kőszemélyek* 33, 336-43(1930).—Oats were grown in exptl containers filled with an alk (limed) soil and with a weakly acid one. Fertilized alk soils showed evidence of translocation of some of the added P_2O_5 . The greatest amt. of P_2O_5 was consumed in the horizon 10 to 20 cm in depth. Unfertilized soils showed no evidence of translocation of the native P_2O_5 . These soils showed the highest content of citrate-sol P_2O_5 in the horizon 10 to 20 cm in depth. As young plants oats require a large amt. of P_2O_5 .

S. S. DE TIVALLY

Laws of solubility of phosphates and potash in mineral soils. R. GANSEN. *Mit Lab Preuss Geol Landesamt* No 9, 1929.—The influence of the Al_2O_3 content of soils on the soly. of K and phosphates is indirect. A high Al_2O_3 content in soil usually corresponds with a high proportion of base-exchanging silicates which tend to "fix" K, and also with relatively large amts. of Al_2O_3 gel. with which phosphate combines. Humus increases the soly. of phosphates in soil by decreasing the proportion of active Al_2O_3 .

B. C. A.

The fineness of grinding phosphates for podsolized soils. N. D. SMIRNOV. *Udobrenie i Urozhai (Fertilizers and Crops)* 2, 765-9(1930).—Eight different raw phosphates, among which were also Florida and Tennessee phosphates, were ground up to various degrees of fineness, and their effects were studied in pot expts. on podsolized soils. It was found that the fractions with a fineness of 0.01 mm. and lower gave 75% and 100% effects, resp., in comparison with acid phosphate. Iral and Florida phosphates gave but slight effects, whereas Tennessee phosphates and apatite gave no effects even with the finest fractions.

J. S. JOFFE

The influence of liming chernozem on the yield and chemical composition of wheat. V. N. KHARCHUKOV. *Udobrenie i Urozhai (Fertilizers and Crops)* 2, 751-9(1930).—The results of 5 years' liming expts. on chernozem with N, P or a combination of the 2 are summarized as follows. At a moisture content equal to 40% of the total moisture-holding capacity limestone alone increased at times the yield of wheat, also in combination with P there was an increase. The neg. results were more frequent than the pos. At 60% of the total moisture-holding capacity there was an increase on the unfertilized plots as well as on the plots fertilized with either N or P. With N and P together liming gave neg. results. The N content of the wheat increased whenever the lime depressed the yield, but it decreased the N content whenever the yield increased. Caustic lime was more effective than limestone in connection with monosodium phosphate when the amt. of phosphate did not exceed the amt. added with manure. With $(NH_4)_2SO_4$ the caustic lime produced very little effect. Analyses of the wheat show that caustic lime increased the N and P content of the grain, indicating that the $Ca(OH)_2$ is capable to mobilize the N and P content of the soil.

J. S. JOFFE

Nitrogen transformation in cultivated soil. F. HANSEN. *Tids Plantet* 35, 713 (1929).— NH_4 salts applied to soil at the end of April or early in May are completely converted into nitrate by the end of June. After plowing in liquid manure or farmyard manure, the NH_4 contents are converted into protein like substances which by mid June are being rapidly mineralized at a rate which exceeds the assimilation of NH_4 by the microorganisms. Nitrification of the NH_4 derived from well rotted manure is more rapid than that of ammoniacal fertilizers. The effects of various nitrogenous fertilizers on soil reaction are discussed.

B. C. A.

Soil and cane composition in relation to Lahaina failure at Waipio Substation. F. HANSEN. *Repts Assoc Hawaii Sugar Tech* 6, 33-7(1927), *Expt Sta Record* 63, 148.—Many factors are causally active in the so-called Lahaina disease, or failure of root development, in the cane of that name. Chem. influences supposedly assocd with growth failure are (1) high salt contents in the soil soln., (2) sol. salts of Fe and Al in the upland acid soils and (3) lack of availability in one or more of the major plant foods. But Lahaina failed in some cases where soil conditions were to all appearances normal, and in order to obtain information regarding this failure in good soils, exptl. plots were planted at the Waipio Substation, the results from which are tabulated and discussed. The failure of Lahaina at Waipio is not assocd with an injurious salt content in the soil, harmful acidity or a shortage of the major nutrients. Studies of the

comprn of the poor Lahaina cane from Waipio contrasted with good Lahaina from Maui showed a lower percentage of potash, a higher content of SiO_2 and a higher Mg/CaO ratio in the Waipio canes. Spectroscopic exams of the ash of poor Lahaina cane from Waipio compared with the ash of the cane from Maui showed the presence of B and Pb in practically all of the latter and in but few cases of canes from the former station. The effect of these rarer elements is being tested at Waipio, as no conclusions appear safe at this time regarding their definite function in plant nutrition. G. C.

Sulfate sulfur in certain types of soils. H. STREMMER. *Chem. Erde* 5, 254-9 (1930). —Peaty soils from the neighborhood of Danzig contain small amts of sulfates, free H_2SO_4 and sulfide S, which vary at different depths and in different seasons of the year (cf. Schroedter, *C. A.* 23, 5449). B. C. A.

The role of peat in increasing the absorption capacity and buffer properties of soils. D. V. DRUZHIKIN. *Udobrenie i Urozhas (Fertilizers and Crops)* 2, 645-50 (1930). —From a series of pot expts. where peat, and with Ca was used with or without fertilizers, with or without CaCO_3 on podzolized soils it was found that peat increases the buffer capacity and thus the adsorption capacity of the soils. Especially is this important in connection with physiologically acid fertilizers. J. S. JORRE.

[Relation among] soil, fertilizers and chemical composition of plants. D. J. HISSINK. *Chem. Weekblad* 27, 529-33 (1930). —Examples are given to show that while the proportions of mineral constituents in the plant depend primarily on the plant itself, they are also dependent on the minerals present in the soil or added as fertilizers. Analysis of the plant and the soil cannot, however, suffice to det. the nature and amt of fertilizer which will produce the best crop. B. C. A.

Fertilizer inspections. Report for 1930. W. C. ROBERTSON. *J. Dept. Agr. Victoria* 29, 136-40 (1931). —Analyses of official inspection samples of fertilizers marketed in Victoria are tabulated. K. D. JACON.

The results of a five-year field experiment on the influence of various fertilizers on the yield and the reaction of strongly acid-exchangeable sandy soils. H. RÖSSLER AND L. SCHMITT. *Landw. Vers.-Sta.* 111, 293-326 (1931). —KCl, kainite and K_2SO_4 were used over a period of 5 years without any increase in soil acidity (NH_4^+SO_4 should not be used on sandy soils which are readily subject to acid exchange without previous application of lime). Of the alk. reacting fertilizers, CaCN_2 exerts the greatest neutralizing action. K_2CO_3 has a harmful effect on the soil acidity. J. R. ff.

Action and value of hot-fermented stall manure. F. LÖNNIS. *Z. Pflanzenernähr. Düngung u. Bodenk.* 9B, 268-72 (1930), cf. *C. A.* 24, 1924. —A reply to recent criticisms (Gerlach, *C. A.* 24, 3882; Lhrenberg, *C. A.* 24, 5099). B. C. A.

The composition of manure and chemical manures. L. BRÉTONNIÈRE. *J. agr. prat.* 55, 269-70 (1931). —A comparison of the analyses of manures made in 1926 and 1931 shows that the amts. of N, P_2O_5 and K_2O are greater in the samples taken in 1931, while there is a decrease in the CaO and MgO content. The decrease in CaO and MgO appears to be due to the H_2O used and the methods of wetting the manure piles. The N, P_2O_5 and K_2O increases seem to be brought about by the use of complementary fertilizers on the soil. J. R. ADAMS.

The decomposition of organic manures. III. The decomposition of green manures. SIGERU OSUGI AND TARO GOTO. *J. Agr. Chem. Soc. Japan* 7, 138-56 (1931). —The quantities of moisture, C and N of air dried green manures, vetches, winter vetches, serradella, *Lespedeza bicolor* Turcz. and oats were estd. for the various stages of growth. The C/N ratio is generally larger according to the growth stage of the plants. The decompn. of these green manures in the field proceeds more readily for younger plants than for older ones. The less the C/N ratio is, the easier the decompn. is. There is the same relation under paddy-field conditions, but the decompn. is generally slow. During the growth of these plants, protein and water sol. N compds. decrease while hemicellulose, cellulose and pentosans increase markedly. The changes in the other constituents are not remarkable or regular. Therefore, the C/N ratio becomes larger for the older plants. Y. KIHARA.

The practical preparation of artificial manure. DELUCQ. *J. agr. prat.* 55, 238-9 (1931). —The straw is spread on the field until there has been sufficient ram to soak it thoroughly. It is then mixed with cyanamide, urea or $(\text{NH}_4)_2\text{SO}_4$ and CaCO_3 , and stacked on a bed of natural manure. After standing a month this artificial manure is ready to be spread on the fields. J. R. ADAMS.

The decomposition of the green parts of lupine in the soil. Z. P. USPENSESKAYA. *Udobrenie i Urozhas (Fertilizers and Crops)* 2, 635-8 (1930). —Green lupine in the blooming stage was added to 5 kg. of soil at the rate of 100 g. and 200 g., which was equal to 50 and 10 g. of N. Another series was prepd. to which 0.5% CaCO_3 was added.

Pptd. phosphate and raw phosphates were used as the source of P. The plant used was millet. The soil was a podsolized loam. The course of the decompn of lupine in soil was followed simultaneously with the pot expts. The 100-g application of lupine gave just as good results as those obtained with $(\text{NH}_4)_2\text{SO}_4$, the 200-g application gave a higher yield. The lime with the 100-g application gave just as good results as the 200-g application alone. Lupine with $(\text{NH}_4)_2\text{SO}_4$ gave the highest yield. The amt of org matter increased, but the increases with the 100- and 200 g applications were the same. The pH of the ext decreased because of formation of the acids. The amt. of absorbed NH_3 increased more in the soils in which lupine was used than in the soils which received $(\text{NH}_4)_2\text{SO}_4$. Composts of lupine at the rate of 5 and 10 g of lupine for 200 g of soils were prepd. These were sampled after 10, 20, 35, 60 and 85 days. The reaction became more alk. on account of NH_3 formation. The nitrate content in the first period decreased with or without addns of lime. From the third period on the nitrates appeared in the soils without lime. With the lime, nitrification was decreased. In all cases there were considerable amts of NH_3 , especially with the lime. With the increase in nitrate content the amt of absorbed NH_3 decreased.

J. S. JOFFE

The effect of the main forms of nitrogen fertilizers. V. P. BUL'SKII. *Udobreniia i Urozhai (Fertilizers and Crops)* 2, 600-13 (1930).—From a series of expts for 3 yrs with various forms of N on the different soil zones with several crops the following conclusions have been made: (1) In the podsol zone on the forest gray soils and on the degraded chernozem N gives large yield increases. (2) In the southern chernozems—pre-Caucasian and Azov types—N, as well as a complete fertilizer does not give any appreciable increases. The deep chernozems occupy an intermediate position. (3) A definite relation has been noted between the physiol reaction of the N sources and the K and P fertilizer and yield increase in the various soil zones. In the podsol zone the basic N salts added to an acid K-P fertilizer gave better results than the physiologically neutral or acid N salts. A similar tendency was observed on the forest gray soils and degraded chernozems. In the southern chernozems the acid N salts were better with the acid K-P fertilizer. (4) The physiologically neutral salts— CaCN_2 , $(\text{NH}_4)_2\text{CO}_3$ and urea—gave considerable increases in yield, but at times they are the poorest sources of N. This is due to the instability of the first 2 and to the methods of applying them. (5) NaNO_3 and $\text{Ca}(\text{NO}_3)_2$ proved to be the best sources of N. Whenever a physiologically acid salt was required, the NaNO_3 and $\text{Ca}(\text{NO}_3)_2$ were not as good as the others. By careful manipulation the CaCN_2 , $(\text{NH}_4)_2\text{CO}_3$, and urea can be used to just as good advantage as the other N fertilizers. No objection was found to the use of $(\text{NH}_4)_2\text{SO}_4$, especially when lime is to be used. NH_4NO_3 , when added to an acid fertilizer, was not so effective as the other forms of N, but with a neutral fertilizer it was as good as any av. N carrier. NH_4Cl proved to be inferior except on the chernozems. Almost all forms of N seem to decrease the starch content of potatoes.

J. S. JOFFE

Organic industrial wastes as nitrogen fertilizers. S. P. GUSEV. *Udobreniia i Urozhai (Fertilizers and Crops)* 2, 614-32 (1930).—Stomach contents from slaughter-house wastes contg 0.18% N, 0.86% P_2O_5 and 0.55% CaO and stomach contents with 1.27% N, 1.51% P_2O_5 and 1.04% CaO were tested in pot and field expts. The results with potatoes and oats indicate that these wastes can be utilized to advantage, but since their keeping qualities are poor and the low fertility value prohibits long hauls they can be used only in the vicinity of the slaughter houses. Mineralization expts.—accumulation of nitrates as an index—show that these substances are slowly mineralized and the addn of lime increases the amt and speed of mineralization. Waste products from the leather industry have a high N content—from 8.82% in leather shavings after tanning to 5.83% in the green hides. But the N is not available. Composting these for 5 months, even with CaCO_3 , gave no available N. Vegetation expts in pots and fields showed that leather waste decreased the yield of oats. These products have to be treated with either acids or alkali before they can be used as a N source. From the data available at the Inst for Fertilizers, the alkali treatment is preferable. Other wastes from the leather industry, like the inside trimmings from the raw hide, which contain about 1.86% N were tried out and proved to be useful, but they can be used locally only. Horn shavings contg 13.97% N, 0.124% P_2O_5 and 0.55% CaO and French burned horn contg 14% N were tried in pot expts. with oats. The results show that 50% of the N was utilized. Various tobacco dusts contg about 2.35% N were used in pot expts with oats. It was found that the N is rapidly mineralized, and in the early stages accumulations of NH_3 were injurious. Wastes from intestines when dried gave a material contg 11-12% N which is very readily available. The

pulp from the fruits in the fruit juice industry, contg up to 1.5% N, proved to be of no value except the residues from cherry juices, which was somewhat beneficial for oats. Bouillon from steamed dead carcasses with a N content of 1.9% proved to be readily available and can be used locally. Blood meal obtained from blood, after the albumin has been sepd out, with 10% N content, is readily available. Ground peat, contg 1.48% N, on oats was not effective because of its slow availability. Sewage waste was tested in pot expts., it was found to be of some value, but its bulkiness is a hindrance in its utilization. Feathers and down proved to be easily mineralized. They contain 6.69% N. The large vol. these materials occupy makes them hard to handle. Wastes from fats used in margarine contain 1.73% N, but they gave neg results. Shells from sunflowers proved to be injurious, producing denitrification reactions. Wastes from the sausage industry gave very good results. Fish scrap contg 9.92% N and 7.26% P_2O_5 gave very good results with oats. Wastes from the wool industry gave fairly good results in vegetation expts. with oats. Waste from the silk industry contg 10.44% N and 2.22% P_2O_5 are easily mineralized and give good results. J. S. J.

Action of various nitrogenous fertilizers on acid soils. D. MEYER, *Z. Pflanzenernähr. Düngung u. Bodenkd.* 9B, 474 (1931). The injurious effects of cyanamide on acid soils previously referred to (cf. Meyer and Obst, *C. A.* 24, 809b) are limited mainly to acid soils poor in humus and of low biol. activity. B. C. A.

Nitrogen fertilizer for the potato. JEAN DETERRÉ, *Engrais* 46, 207-11 (1931).—A discussion of the problems arising in the fertilizing of potatoes. The proper fertilizer is one with a phospho-potash base and with N in the NH_4 form. J. R. ADAMS.

Nitrogen top-dressing of wheat. Experiment with various forms of fertilizer. A. Y. MONTGOMERY, *New Zealand J. Agr.* 42, 185-92 (1931).— $(NH_4)_2SO_4$ was the most satisfactory N fertilizer for the top-dressing of wheat, followed in order by $NaNO_3$, urea, Ca cyanamide and Calnitro. Because of its poor phys. condition, it was difficult to apply small quantities of urea through a drill while it was impossible to drill Calnitro. The irritating effects of Ca cyanamide on the eyes and skin of the operator constitute a serious drawback to its use as a fertilizer. The yield of flour percentage of protein, and the baking qualities of the flour were not affected to a significant extent by the different N fertilizers. K. D. JACOB.

The effect of nitrogen fertilizers on the yield and quality of sugar beets. A. G. MIKHAILOVSKII, *Udobreniia i Urozhai (Fertilizers and Crops)* 2, 633-8 (1930).—Urea, Ca cyanamide, Chili salt-peter (NH_4) $_2$ SO_4 , Leunaspalt-peter and NH_4Cl , together with a mixt. of P and K, were tried out on degraded chernozem soils for beets. $(NH_4)_2SO_4$ proved to be most effective, followed by urea, NH_4Cl and Leunaspalt-peter. $NaNO_3$ and $CaCN_2$ gave neg results. The highest sugar content was obtained with $NaNO_3$, the sugar content was increased 1.5%. This is followed by urea, Leunaspalt-peter and NH_4Cl decreased the sugar content in the beets. The amt. of "injurious" N which accumulates in sugar beets is affected variously by the different N sources. $CaCN_2$ gave the highest value, followed by $NaNO_3$. In expts. without P the yield increase for N was only 5%, whereas with P it became 18%. J. S. JORIS.

The effect of mineral fertilizers and manure upon the quality of sugar beets. I. A. SIRACHENKO, *Vustk Zefskis Traktors Prom.* 10, 477-90 (1930).—With repeated application of mineral fertilizers the pH of the soil decreases. Easily sol. compds. in the upper layers move downward and therefore increase the fertility of the lower layers. The mineral fertilizers, compared with manure, decrease the sucrose and maltose content in the beet. The development of the beet is better in the first part of the summer with manure and in the second part with mineral fertilizers. V. L. BAIKOV.

Fertilizing action of magnesia. CH. BRIQUET AND EDG. JORIS, *Compt. rend. acad. agr. France* 17, 201 (1931).—Pot tests with white mustard show that the addn. of a fertilizer contg MgO to a soil low in total and assimilable MgO produces a marked increase in the dry wt. of the crop and in the MgO content of the dried material. The MgO also appears to favor the absorption of P_2O_5 , but has little effect on the absorption of N. Similar expts. with maize indicate an increased yield and quantity of MgO in the dried plant, but the MgO has no apparent influence on the absorption of P_2O_5 or N. The addn. of a MgO -bearing fertilizer to a field planted in sugar beets gave a marked increase in the yield. J. R. ADAMS.

The action of chemical fertilizers on native meadows. R. BONNICEL, *J. agr. prat.* 55, 250-1 (1931).—A preliminary report on the effect of fertilizers on soil plots of various geological formations shows that the most efficacious fertilizer is not necessarily the one contg the elements that are deficient in the soil. The natural richness of some of the soils in K_2O does not impede the action of added K_2O . J. R. ADAMS.

A study of the composition of pineapple plants at various stages of growth as in-

finenced by different types of fertilization. J. M. HORNER. *Hawaii. Pineapple Cannery's Sta Bull* 13, 42 pp (1930). *Expt Sta Record* 63, 641.—The pineapple crop removes more potash from the soil than any other mineral constituent, under the conditions of the expt an av. of 1450 lbs per acre of K_2O as compared with 491 lbs of N and 111 lbs of P_2O_5 . The rate of growth was very similar on all plats, being slightly but not significantly greater on the potash areas. Potash constituted about 37, 32 and 45%, resp., of the total ash in the leaves, stems and fruit. There was a very close similarity between the potash and total-ash curves. The greatest content of total and reducing sugars was reached at fruiting time, being apparently correlated with the age of the plant. Starch content was highly variable, starch continuing to accumulate on some plats even during the time of fruit formation when large quantities of sugar were being produced. The ratio between N and Ca was observed to be fairly constant. N stimulated growth for a short time but did not appear to have an enduring influence. P did not cause as large growth as did the other nutrients, and in excess quantities actually depressed growth. It recommends the shredding of pineapple waste and its return to the soil. G. G.

The availability of phosphatic fertilizers as shown by an examination of the soil solution and of plant growth. A. W. GAREVILL. *J Agr Sci* 20, 559-72 (1930).—In pot expts conducted with an acid soil deficient in P_2O_5 , liming with $CaCO_3$ at the rate of half the CaO requirement increased the concn of P_2O_5 in the soil soln from about 0.7 ppm to about 1 ppm, and at the rate of twice the CaO requirement to over 2 ppm. Phosphates depressed the concn on the lightly limed soils and had variable effects on those heavily limed. The superphosphate treated soils showed rather lower concns than those treated with slag. Cropping raised the concn on the lightly limed treatments, but on the heavily limed reduced it on the control and had a variable effect on the phosphate treatments. The concn of P_2O_5 in the untreated soil was practically constant. After treatment it fluctuated considerably, but tended to become more constant later. There was no indication of relationship between the concn and the soil-moisture content. The heavily, compared with the lightly, limed treatments produced more extensive tillering and higher yields. Phosphates increased the yields, particularly of grain, but had no effect on tillering. Slag and superphosphate were equally effective on the heavily limed soil, but the latter was the more effective on the lightly limed. No correlation was apparent between crop growth and the concn of P_2O_5 in the soil soln. The possibility of plant roots taking up phosphate directly is suggested. In further pot expts higher yields were obtained from freshly treated soils than from soils treated a year before sowing on the lightly limed treatments, while the reverse was the case with the heavily limed. On the soils treated the previous year the heavily limed treatments gave higher yields than the lightly limed. On the freshly treated soils the 2 gave similar yields with early sowings. Slag and superphosphate gave similar results throughout, and their application was more effective on the freshly treated soils than on those treated the previous year. P. R. DAWSON.

Quantitative analysis of phosphoric acid. VI. Determination of phosphoric acid in a phosphate fertilizer. M. ISHIBASHI. *Mem Coll Sci. Kyoto Imp Univ* 13A, 291-301 (1930), cf *C A* 23, 4645, 4907, 25, 2663.—I's improved volumetric phosphomolybdate method is especially suitable for the detn of small quantities of H_3PO_4 . For the extn of free H_3PO_4 , $MgCO_3$ is more satisfactory than either alc. or ether. B. C. A.

Experiments with increasing dressings of phosphoric acid on Riesling and Burgundy vines. C. DREYSPRING AND H. KERTH. *Superphosphate* 4, 93-103 (1931).—The height, no. of knnts, internodes and thickness of the wood of young grape vines were in direct relation to each other, and to the amt of phosphate fertilizer applied, up to 150 kg of P_2O_5 per hectare. The total wt. and percentage of P_2O_5 in the wood of the vines were also directly related to the amt of phosphate applied. K. D. J.

The decrease in citrate-soluble phosphoric acid in Rhenania phosphate. K. SCHARER. *Landw Vers Sta* 111, 110 (1930).—Fifteen samples of Rhenania phosphate were kept for a year in cork stoppered vessels at room temp. All showed a decrease in citrate-sol P_2O_5 . But only 5 showed losses (0.3-0.54%) greater than the exptl error. The decrease is due to the action of moisture. JOHN R. HILL.

The production of precipitated phosphate from Aktyubinsk raw phosphates. K. K. APUSHKIN AND L. M. GUREVICH. *Udobrenia i Urozhai (Fertilizers and Crops)* 2, 740-8 (1930).—Increasing the concn of H_2SO_4 from 15 to 25% decreased the coeff of decompn of raw phosphate from 95 to 93.1%, but the concn of P_2O_5 in the ext. increased from 5.86 to 10.67%. At a 25% concn of H_2SO_4 (the stoichiometric quantity) the pulp was so thick that it barely flowed on the filter. By dilg the acid with wash water

the viscosity of the pulp increased, and it worked nicely even with 30% H_2SO_4 . Even with 10% H_2SO_4 below the stoichiometric quantity the coeff. of decompn. was as good as with optimum concn. With a 15% concn. of H_2SO_4 with 10% below the stoichiometric the time factor is of no importance. By using a 4% concn. of H_2PO_4 to dil. the 15% H_2SO_4 , the coeff. of decompn. decreased 2%, but the amt. of P_2O_5 in the ext. increased 2 to 3%. With 8% P_2O_5 , the coeff. decreased slightly, and the amt. of P_2O_5 increased up to 15.78%. The most economical procedure for the production of pptd. phosphate was thus 15% concn. of H_2SO_4 diluted with water, the time of contact being 40 min. A 5.2% soln. of milk of lime gave from the Aktyubinsk exts. a ppt. of 37.5 to 40.8% P_2O_5 , of which 90% is citrate sol. if dried at 80%. J. S. JOFFE

The new complex ammonium phosphates and the raw materials for their manufacture. UGO ORLANDI *Industria Chimica* 6, 18-22(1931)—The 3 complex phosphates $\text{NaH}_2\text{N}_2\text{H}_4\text{P}_2\text{O}_6$, $\text{KH}_2\text{N}_2\text{H}_4\text{P}_2\text{O}_6$, and $\text{MgN}_2\text{H}_4\text{P}_2\text{O}_6$ have become very important in the fertilizer industry since the development of the manuf. of synthetic NH_3 . The raw materials used and the manuf. of the above 3 fertilizers are described, and their role as fertilizers is noted. A. W. CONTIERI

Manufacture and use of insecticides and fungicides. V. Organic compounds. ARTHUR S. CARLOS. *Fertilizer, Feeding Stuffs and Farm Supplies J.* 16, 222-3(1931); cf. *C. A.* 25, 763—The manuf. and use of nicotine, pyridine, hellebore and pyrethrum preps. are briefly discussed. K. D. JACOB

Fumigants. C. L. WILLIAMS. *U. S. Pub. Health Repts.* 46, 1013-31(1931)—This is chiefly a discussion of the use of HCN . Other fumigants are briefly mentioned. J. A. KENNEDY

The control of loose smut in wheat. D. B. ADAM AND R. T. M. PEARSCOTT. *J. Dept. Agr. Victoria* 29, 141-5(1931)—Practically complete control of loose smut of wheat (*Ustilago tritici*) was obtained by soaking the seed in water at 120°F. for 1 hr. and 40 min. The treatment had a depressing effect on the germination of the seed, the effect increasing with the length of time of treatment, but the total yield of the crop from the treated seed was not affected. K. D. JACOB

"Target spot" of tomato seedlings. S. FISH. *J. Dept. Agr. Victoria* 29, 132-3(1931)—The percentage of "target spot" infection, caused by the fungus, *Macrosporium solani*, on tomato seedlings was reduced from 97 to 17.5 by spraying with Bordeaux mixt., prepd. in the proportions of 6 lb. CuSO_4 and 4 lb. CaO to 40 gallons water. K. D. JACOB

Heart rot of fodder and sugar beets. E. GAUMANN. *Landw. Jahrb. Schweiz* 44, 143-50(1930)—The disease occurs exclusively on alk. or neutral soils. When the pH falls below 7.8, beet cultivation is unprofitable because of the disease. Susceptible varieties contract the disease at pH 7-7.2, while the resistant varieties remain healthy at pH 7.6. Where soil conditions are critical the use of resistant varieties may serve as control of the disease. Omission of N from the fertilizer caused an increase in the infection. GEORGE R. GREENHANK

Molds utilized to make sewage sludge odorless (WIGLEY) 14. Determination of pH by means of an Sb oxide electrode (DI GLORIA) 7. Some properties of commercial nitrates of lime (MATIGNOV) 18. The influence of frequency of cutting on the productivity, botanical and chemical composition and the nutritive value of "natural" pastures in Southern Australia (DAVIES, SM) 11D. Use of HCN gas (Fr. pats. 697,023-4) 18.

Journal of the Imperial Agricultural Experiment Station (Japan) (*New journal*). Published by the Imperial Agricultural Experiment Station, Takinogawamachi, near Tokyo, Japan. Vol. 1, No. 1, appeared in 1930.

Memoirs of the Tottori Agricultural College (*New journal*). Published by the Agricultural College, Tottori, Japan. Vol. 1, No. 1, appeared in May, 1930.

WAKSMAN, SELMAN A., AND STARKET, ROBERT L. The Soil and the Microbe. New York. John Wiley & Sons, Inc. \$3.50

Fertilizer. HANS BRENEK (to Rhodania Verein Chemischer Fabriken A.-G.). U. S. 1,799,892 April 7. A mixt. of raw phosphate, silica and alkali metal carbonate is heated, in the absence of steam, to about 1200°. The mixt. used contains about such amounts of P_2O_5 , CaO and SiO_2 , and at least such an amount of alkali metal oxide as would be present in a mixt. of $\text{Ca}_2\text{Na}_2\text{P}_2\text{O}_7$ and Ca_2SiO_4 having the same content of P_2O_5 and CaO as the raw phosphate employed.

Fertilizers. PAUL ASKENASY, ALFRED STERN, FRIEDRICH NESSLER and ANDREAS

VON KRISLER Ger. 516,970, Dec 23, 1928 A mixt. of $\text{NiH}_2\text{I}_2\text{PO}_4$ and KfH_2PO_4 for use as a fertilizer is prepd. by treating an aq. soln. of $\text{NiH}_2\text{I}_2\text{PO}_4$ with a suitable amt. of KCl . The soln. is preferably hot. Examples give the proportions of KCl $\text{NiH}_2\text{I}_2\text{PO}_4$, as 1:1 and 1:2 A 94.6% yield of $\text{NiH}_2\text{I}_2\text{PO}_4$, KfH_2PO_4 , is given

Fertilizers. GASVERARBEITUNGSGES M B H Fr 697,680, June 17, 1930 Fertilizers are kept stable by adding salts such as $\text{Ca}(\text{NO}_3)_2$ or CaCl_2 which have hygroscopic properties in amt. sufficient to oppose any change of moisture content of the fertilizer

Fertilizers. I G FARBENIND A-G Fr 697,918, June 26, 1930 The double salts $\text{K}_2\text{H}_2\text{SO}_4$, NO_2 or the double salt $(\text{NH}_4)_2\text{H}_2\text{SO}_4$, NO_2 or the mixts of salts obtained, as the double salts mentioned, by directing nitrous vapors into a H_2SO_4 and HNO_3 soln. of K or NH_4 salts are mixed with appropriate substances for forming compd. fertilizers Cf C A 25, 2236

Fertilizers. SOC CHIM DE LA GRANDE PAROISSE (AZOTE ET PRODUITS CHIMIQUES). Fr 697,966, Oct 1, 1929 A compd fertilizer is made by the action on natural phosphate of gases coming from the oxidation of NH_3 without a preliminary transformation into HNO_3 . A mud of phosphate suspended in water may be used Cf C A 25, 1026

Fertilizers. KUNSTDÜNGER PATENT-VERWERTUNG A G Fr 698,155, June 27, 1930 In the treatment of phosphatic rock by HNO_3 and pptn of CaSO_4 by K_2SO_4 , the reaction products are dild by the addn of a soln contg phosphoric acid or an NH_4 salt or both before pptn of CaSO_4 . A mixt. of phosphoric acid and CaSO_4 obtained by dissolving phosphatic rock in H_2SO_4 may be used as addn

Fertilizers. ALFRED MENZEL Fr 698,424, July 5, 1930 Granular or fragmented masses such as fertilizers contg easily volatilized or decomposed substances are stabilized by coating the grains with a more stable substance which is volatilized and condensed on the grains NH_4HCO_3 may be coated with NH_4Cl , urea or $(\text{NH}_4)_2\text{SO}_4$

Fertilizer. KIYOSI OKADA and OSAXA ALKALI TOTI KABUSHIKI KAISHA Japan 80,510, Dec 11, 1930 Extract of Ca_3N_2 is mixed with H_2PO_4 , K_2SO_4 or $(\text{NH}_4)_2\text{SO}_4$ is added, CaSO_4 filtered off CaSO_4 and the soln. concd and mixed with NH_4 , K_2CO_3 , $\text{Ca}_3(\text{PO}_4)_2$ to neutralize the remaining free H_2PO_4

Moisture-proof fertilizer. KÔTA OSADA and NIPPON KALI KOGYO KABUSHIKI KAISHA Japan, 80,509, Dec 10, 1930 The fertilizer is a mixt. of $(\text{NH}_4)_2\text{SO}_4$ 40, Ca perphosphate 46.5, KCl 13.5, Ca_3N_2 2 and H_2O 5-10 parts

Fertilizer from distillery and sugar factory residues. ZDENKO METEL (to Soc. anon Selbi, Soc d exploitation de licences de brevets industriels) U S 1,709,176, April 7. In prep. a dry and non hygroscopic fertilizer from ale distillery or sugar factory residue and superphosphate, the manuf. is effected in 2 sep. stages at different temps, the residue being heated in the first stage to about 70° with a quantity of superphosphate the wt. of which is about equal to that of the residue, while in the second stage the temp. is raised to about 130° to complete the expulsion of water and org. acids set free in the first stage.

Ammonium potassium phosphate fertilizer. KASEI KUN and KATAKURA BEIKOKU HIROYO KABUSHIKI KAISHA Japan 90,174, Feb 6, 1931 Excess of H_2SO_4 is mixed with Ca perphosphate to make a soln contg free H_2SO_4 and H_2PO_4 in the ratio of 2:1. Ca_3N_2 , $(\text{NH}_4)_2\text{SO}_4$, KCl and org. matter such as soy bean cake are added. The whole is kept for about 24 hrs. in a brick chamber to complete the chem. reaction. The product is sol. in H_2O and contains no cyanamide

Insecticide and fungicide. FREDERICK W SULLIVAN, JR., and ELMER W ADAMS (to Standard Oil Co of Ind.). U S 1,800,114, April 7 A compn. capable of being dispersed in water to form a relatively stable horticultural insecticidal and fungicidal emulsion comprises mineral oil, an As compd. such as Pb arsenate, an oil sol. Cu rosni soap, a non soap colloid such as glue and mineral oil sulfonic soap

Fungicides and insecticides. FRITZ LEIBERANDT. Fr. 697,638, June 18, 1930 Compds contg Cu and As which are valuable fungicides and insecticides are prepd. by transforming Schweinfurt green (a mixt. of arsenite and acetate of Cu) with salts formed from a strong base and an acid which is weaker than AcOH and preferably stronger than arsenous acid, e. g., a fatty acid of high mol. wt., the amt. of salt being such that the 2AcOH residues of the Schweinfurt green are changed. An example is given using Na palmitate

Insecticide. YUZO TUBAKIMOTO and DAICHI KÔGYÔ SEIYAKU KABUSHIKI KAISHA. Japan 90,229, Feb 10, 1931. The insecticide is a mixt. of sericia, Cu soap, gum arabic, agar agar and PbHAsO_4

Protecting plants against insects, etc. RICHARD SCHLÖGL. Austrian 122,023,

Nov. 15, 1930 Use is made of the sulfurous waste waters from the manuf. of films, artificial silk, etc., from viscose

Anticryptogamic products. SOC. SELOXYDE Fr 698,604, Oct. 8, 1929 An anticryptogamic paste is obtained by mixing a hydrated oxychloride of Cu and a compd such as a base, a carbonate or a basic salt capable of displacing HCl from this oxychloride

Seed disinfectant. MORRIS S KHARASCH and MAX ENGELMANN (to E. I. du Pont de Nemours & Co.) Can 309,904, Mar 31, 1931 Org mercury compds are produced by the reaction of $HgCl_2$ with $PbEt_4$ in the presence of a hydrocarbon solvent. These compds are used in the prepn. of seed disinfectants. Cf C. A. 24, 4579

Seed disinfectant. MORRIS S KHARASCH, MAX ENGELMANN and WENDELL H. TISDALE (to E. I. du Pont de Nemours & Co.) Can 309,905, Mar. 31, 1931 A seed disinfectant in dust form comprises 1 $HgCl_2$, tolanite, activated carbon and lime, the total amt. of diluent being not less than 90% by wt. Cf C. A. 25, 2237.

Organic mercury compounds and dust disinfectants containing these compounds. MORRIS S KHARASCH and MAX ENGELMANN (to E. I. du Pont de Nemours & Co.) Can. 309,903, Mar 31, 1931 A seed disinfectant comprises the product of a reaction between $HgCl_2$ and $PbEt_4$ obtained directly in dust form in the presence of an inert solid diluent.

16—THE FERMENTATION INDUSTRIES

C. V. FREY

Science and the fermentation industries. ARTHUR R. LING *J Inst Brewing* 36, 446-55(1930)—An outline is given of the rise of scientific education along brewing lines and the need for its continuation.

Progress and development of the fermentation industries. A. FERNBACH. *Chimie & Industrie* 25, 570-6(1931)—A brief review

Influence of pH on alcohol recovery in cane-molasses distilleries. N. E. LAMONT. *Intern. Sugar J* 32, 630(1930). *Sugar Abstracts in Facts About Sugar* 26, 88—In expts. made by the author the following acidities were maintained during fermentation of molasses washes in the wash before entering the vat, and before fermentation commenced, 4.4, on completing the filling of the vat, 4.0, after fermentation, 3.7, and 12 hrs. after fermentation was stopped, 3.6 Ordinarily, the yield of alc. wash was 5.5% (alc. of 42.5 Cartier), but repeated expts. along the lines indicated above gave a yield of 6.5 and over This practice is therefore being continued as standard, once a week the vats are thoroughly washed and disinfected.

Sulfurous and lactic acids in the production of fruit wines. REIDIGER and K. SICHERT *Wein u. Obst* 11, 50-8(1930)—The addn. of lactic acid (2-6 g. per l.) to perry prepd. from pears low in acid is recommended in place of that of H_2SO_4 .

Institute of Brewing research scheme. L. The prediction of extract. L. R. BISHOP *J Inst Brewing* 36, 421-34(1930)—A statistical study made of the analytical data of the Institute research scheme shows, in any one variety of barley, that the N content is in reverse relation to the yield of ext. of the resulting malt. For one type of barley the ext. of the malt can be predicted within close limits by the equation $E = 110.1 - 11.2 N + 0.18 G$, where E = ext. in lb. per quarter of dry malt, N = N % of dry barley and G = wt. in g. of 1000 dry barley kernels. The equation assumes similar modification of the malt, although individual maltsters can adjust the const. (110.1) to their requirements. Also for other varieties of barley, predictions can be made with reasonable accuracy by a change in the const. of the equation.

Adsorption in the brewing industry. J. RAUX. *Brasserie & malterie* 21, 6-12, 22-4(1931)—An address outlining the nature and laws of adsorption and the part it plays in beer making

Steam turbine as power source in brewing. H. HERBRACHT *Wochschr. Brau* 47, 543-54(1930)

The Elmocid process (a new way of disinfection). F. STOCKHAUSEN *Wochschr. Brau* 47, 523-8, 545-9, 561-4, 575-9(1930)—A new process of disinfection was developed by the Elektro-Osmose Gesellschaft. Two products are used "Alkaline-Elmocid" (0.58% NaCl and 0.4% NaOH, pH 13.07) and "Acid Elmocid" (NaCl and HNO_3 , both in 0.1 N concn., pH 1.07). Alc. Elmocid is generally used, whereas Acid Elmocid is preferred for Al tanks and pitch-coated equipment. The efficiency of these products

was carefully investigated, and the reports are favorable in both small scale lab. and large-scale factory expts. The disinfecting action of "Alkaline Lincold" was found to be somewhat greater, or at least more rapid, than that of the acid product. The advantage of Alkaline Lincold over hypochlorite is emphasized by its chem. stability. It can be used over again several times and brought up to its original strength. Conclusion: The process can be used successfully in breweries. S JÓZSA

Fluorescent indicators in acidity determinations of colored wines. Y. VOLMAR AND J. M^o. CLAVIERA. *Anales soc. esp. fis. quim.* 29, 247-51 (1931).—The application of the volumetric method with the fluorescent indicators acridine and umbelliferone in Wood light to exact titration of total acidity in red wines was studied. Umbelliferone at a concn of 4 to 6 drops per 5 cc. of wine dild. to 100 cc. gives a change at exactly $pH = 7$. E. M. SUMMERS

Malt dust, its disturbing effect on fermentation. K. SCHUSTER. *Wochschr. Brau.* 47, 451-3 (1930). The last portions of malt from large concrete silos often showed unsatisfactory fermentation attributed to malt dust. Expts. on a small scale proved this assumption: the fermentation was sluggish at first, the head was undeveloped and coarse in texture. The final attenuation is reached in the vat; the yeast does not settle well and shows a dusty or powdery character. The beer leaving the vat is cloudy; the haze in many cases cannot entirely be removed. S. advised handling and transporting the malt carefully, avoiding rough treatment, to minimize the formation of dust. Too fine grinding will cause similar troubles. S JÓZSA

Determination of liquefying power of malt amylase. L. FLETCHER AND J. B. WESTWOOD. *J. Inst. Brewing* 36, 650-7 (1930).—Investigation of the method of Józsa and Gore (C. A. 24, 1170) shows results to be relative if all conditions are kept constant. Within certain limits the changes taking place during liquefaction comply with Kjeldahl's law of proportionality, but it is necessary to increase the proportion of enzyme to substrate to get near 100% liquefaction, since Józsa and Gore's curve, as well as tests made, shows that with the low ratio of Józsa and Gore, liquefaction comes to a practical standstill at 90%. F. E. J. P. WENNER

Observations of some malting and brewing trials with a six-rowed winter barley. H. LLOYD HIND. *J. Inst. Brewing* 36, 435-9 (1930).—Growing, malting and brewing tests of a British-grown 6 row barley. F. E. J. P. WENNER

Action of ultra-violet rays on the germination of barley in malting. R. DE PAET. *Ann. Acad. Lincei* [8], 11, 1022-1 (1930).—The action of ultra violet rays on germinating barley accelerates the germination even at low temps (10-15°), prevents mold growth, facilitates the formation of enzymes and increases the vitamin content. Expts. on the large scale show that malt of improved quality is obtained in this way. B. C. A.

Swedish brewing barley, its cultivation, improvement and brewing value. C. WIRTH. *Wochschr. Brau.* 47, 483-6, 495-8, 505-8 (1930).—The principal barley-growing districts are the provinces Schonen and Östergötland and the islands Gotland and Öland. The early ripening barleys are chiefly cultivated, and great importance is attached to purity of type. The details of cultivation are given. These Swedish barleys which are better suited for pale than for dark beers, have a rather high moisture content (17-17½%), and drying is advisable. The samples showed a low protein content and high yield on ext. The worts saccharify well, and the beers obtained show a good head forming power. S JÓZSA

The nitrogen content and "quality" of barley. L. R. BISHOP. *J. Inst. Brewing* 36, 352-69 (1930), cf. C. A. 24, 1929.—A résumé of studies made and suggested required knowledge to lead to a sound, quant. method of assessing barley value. P. J. F. W.

Barley husk. H. LEWIS. *Wochschr. Brau.* 47, 637-61, 671-5 (1930).—The aq. extn. of various constituents from barley husk was studied under different conditions. Detailed analysis of the husk gave fat 2.22, pentosans 37.95, protein 5.9 and ash 5.9. The ash analysis showed P_2O_5 , 5.70, SiO_2 , 64.91, Ca 2.77 and Mg 2.06%. The results showed that the acids might be removed from the husk by a suitable alk. treatment. In the steep. Barley was subjected to a preliminary steep contg. 0.2% NaOH. The malt was equal to the check, and the pale beer was finer and milder in flavor. L. suggests such steeping for barleys having a rough and coarse husk. It is not necessary, however, when dark malt and beer are processed. S JÓZSA

Modern wine-making methods. J. A. DONROW. *Can. Chem. Met.* 15, 93-4, 97 (1930).—The control of microorganisms, the development of wine yeast-control factors and temp. control are discussed. W. H. BOYNTON

Improvements in the fermentation and maturation of beers. L. NATHAN. *J. Inst. Brewing* 36, 538-60 (1930).—A comparison of the English brewing methods and

the Nathan System, which consists essentially of quick removal of boiled wort from the hops, aeration with sterile air and removal of volatile constituents, chilling and removal of suspended matter, and pitching with pure, culture yeast. The fermentation is regulated by passing finely subdivided CO_2 into the fermenting liquid, purifying the CO_2 after use and subsequently using the CO_2 to remove young flavors from the fermented beer after sepn from the yeast. Beer is then further cooled and satd with CO_2 . The production of a 12% original gravity beer requires 12 days' time.

PETER J. F. WEBER

Studies on tannin. II. Determination of the "tannin-N" and the tannin number as a new method for defining the protein conditions in beer. II. FRANK AND W. RIEDL. *Wochschr. Brau* 47, 491-4, 501-5 511-5(1930), cf. *C. A.* 24, 2830.—The authors found that the amt of nitrogenous matter pptd from beer is a measure of the amt of complex proteins of high mol wt. An analytical method is developed. The term "tannin N" is suggested for the proteins pptd by tannin. The tannin no. of a beer is the fraction (expressed in %) of the total proteins pptd by 1% (tannin no. I) and 2% (tannin no. II) pure tannin solns under exactly specified conditions. The method is similar to the procedure applied heretofore for the detn of coagulable N. The new method gives results which can easily be duplicated. It is immaterial whether the tannin soln is run in rapidly from a pipet or added drop by drop, but the subsequent filtration is facilitated if the tannin soln is run in rapidly and the mixt cooled as prescribed. Low temps are more favorable for the pptn. Various tannin preps will give different results, and comparative estns should be made with the same prepn. The tannin-protein pptn is not in proportion to the tannin concn but decreases in higher concn. The character of the beer is not altered by the tannin pptn, and the tannin no. is not changed by pasteurization. Dark beers show a higher tannin no. than pale ones. 1 sport beers which are especially fit for pasteurization show very low tannin nos.

S. JÓZSA

Thirteenth report on the trial of new varieties of hops, 1929. E. S. SALMON. *J. Inst. Brewing* 36, 533-8(1930).

PETER J. F. WEBER

Some recent ideals on the evaluation of hops. THOMAS K. WALKER. *J. Inst. Brewing* 36, 410-6(1930).—The aspect is given of researches in progress toward correlating analytical data of hops, with the results to be expected in practice.

PETER J. F. WEBER

Analysis of hops. W. WOLLMEYER. *Wochschr. Brau* 47, 521-5(1930).—The method was proposed by W. (*C. A.* 23, 4011) and was adopted, with some modifications, as a standard method by the Analysis Commission of the German Brewing Industries. The sampling and prepn of samples (reduced to a fine uniform powder) are standardized. Moisture is detd at 104-105° for 3 hrs. The exact details of the estn are given.

S. JÓZSA

A colorimetric method for the determination of the soft resins of hops. JOHN M. GUTHRIE AND GEORGE G. PHILIP. *J. Inst. Brewing* 36, 455-9(1930).—A preliminary report. A 0.5% soln of U acetate in MeOH, filtered from insol matter and added to MeOH solns of α - and β hop resins, produces an orange-yellow color of such intensity as to be capable of measurement by comparison with a standard MeOH soln of α -resin. The ratio of the color of the α resin to that of β resin is as 1:1+4.17, which relation is similar to the ratio of concn of α - and β resin of equal antiseptic value, i. e., 1:1+4.3 (cf. Ford and Tait, *C. A.* 20, 2892). As adapted to the soft resin detn of hops, 10 g. of finely minced hops are extd with 100 cc. of MeOH with continuous agitation for 10 min. The supernatant liquid is decanted through a folded filter, and 20 cc. of the filtrate is measured into a separatory funnel, 40 cc. of 0.25 N H_2SO_4 is added and extn of the resins is made with 4 successive portions of 40-50 cc. each of petroleum ether. The petr ether exts are filtered into a 200-cc volumetric flask, the filter paper is washed with petr ether and the vol is made up to 200 cc. Ten cc is then put into a 100-cc volumetric flask and the solvent evapd by immersion for a few min in a water bath at 40°, and the residue is dissolved at once in MeOH. Ten cc of the 0.5% U acetate soln is then added and the vol adjusted to 100 cc. The color is then compared in a Hellige colorimeter with a wedge of a standard color made up with 0.015 g α resin per 100 cc. The colorimeter reading multiplied by 0.15 gives the sum of the α resin and the colorimetric equiv of the β resin in terms of α -resin, expressed as % of the hops. Results are in good general agreement with gravimetric results using the formula $\alpha + (\beta/4.3)$. Where α - and β resin are required separat ly, total soft resins are detd by evapn of a sep portion of the petroleum ether ext, and the % of each resin is calcd from the colorimetric relationship of the 2 resins.

PETER J. F. WEBER

Manufacture of sulfite spirit from waste pulp liquors in Germany during the war period 1914-18 (SEMBRITZKE) 23. Proteolytic enzymes. VII. Peptidases of green malt (SATO) 11A. Coating concrete [fermentation vats] (Fr. pat. 698,562) 20. Fertilizer from distillery residues (U. S. pat. 1,799,176) 15.

Ethyl alcohol. AUGUST UICK Ger 522,380, Aug 14, 1923 In the manuf. of EtOH by fermentation, saccharification of starch is effected with very finely cut up or comminuted malt, as distinct from bruised or crushed malt.

Gluconic acid. SIEGWART HERMANN and PHARMACEUTISCHE WERKE "NORGINE" A. G. Ger 522,147, Feb 20, 1927 Gluconic acid is prepd by fermenting nutrient solns contg saccharose or glucose with a ferment present in a material stated to be known in Japan as "kombucha." Thus, an ext of Russian tea may be treated with 7.5% of sugar and inoculated with kombucha. After standing for 20 days at 15-40°, the mixt. is evapd to remove AcOH, then dild again, neutralized with CaCO₃ and evapd to recover Ca gluconate.

Fermenting cellulosic materials. HERBERT LANGWELL Ger 522,146, Sept 10, 1925. See U. S. 1,602,306 (C. A. 20, 3771)

Apparatus for the production of distilled liquors using a naked gas flame. EMILE GAZAGNE Fr 698,446, July 6, 1930

Distilling flask for determining the acidity of wine, etc. MARTIN VIALATTE (Soc. A. R. N.) Fr 698,630, July 7, 1930

Hopping wort. KARL S. FELIX Ger 522,145, Dec. 24, 1927 The hops, enclosed in a perforated container, are boiled with the wort in known manner, and the container is lifted out and rinsed. The hops are then treated with water under pressure at 110-125°, and the resulting ext. is added to the boiling wort.

Beer. HEINRICH VAN DE SANDT Fr 698,375, July 4, 1930 Vitamins are added to or produced in beer at a stage in the production of the beer when no destruction of the vitamins would take place.

Beer. OTTO BORG Ger 516,978, June 28, 1929 Colored beer is deprived of its bitter taste by treating with O₂ in the cold.

Yeasts. HUGO WINDELSHEIM and FRIEDRICH W. THIELE Fr 698,316, July 2, 1930. The fermenting power of yeasts is increased by adding cholesterol, cholesterol compds or substances resembling cholesterol in a fine state of subdivision. The cholesterol may be in the form of an emulsion with fats. Phosphates or superphosphates may be added at the same time.

Degrading yeast. I. G. FARBENIND A. G. (Otto Ambros and Erwin Walz, inventors), Ger 520,853, Jan 3, 1928 Addn to 517,409 (C. A. 25, 2157). Yeast is subjected to enzymatic degradation, preferably under sterile conditions, until the greater part has become sol in water. The sol portion is then sepd from the insol. residue, without heating, and the soln and the residue are separately worked up. Thus, yeast may be liquified with AcOEt, brought to a hydron concn. of pH 6-7 by addn of dil aq NH₃, and allowed to stand for several days until tyrosine begins to sep. This is removed by means of a fine sieve, and the liquid autolysate is centrifuged to sep the undissolved portion. This contains all the ergosterol, which may be isolated by alk sapon. Leucine, etc., can be isolated in known manner from the liquid portion.

17—PHARMACEUTICAL CHEMISTRY

W. G. KEMERY

General occurrence of alkaloids in Buxaceae. E. MARTIN-SANS Compt. rend. 191, 625-6(1930)—The occurrence of alkaloids (usually situated in morphologically distinct secretory cells) in 9 species of Buxaceae has been demonstrated microchemically, no alkaloid was detected in *Daphniphyllum macropodum*, which is possibly not to be included among the Buxaceae. B. C. A.

Vacuum microsublimation of synthetic drugs. R. EDER and W. HAAS. Mikrochem., Emich Festschr. 1930, 43-82—Various methods available for the sublimation of small quantities of material are described. The types of sublimate obtained with Eder's microsublimation app and with Diepolder's vacuum app are described, and details, including photomicrographs, are given of the results obtained with a no. of synthetic drugs. Four groups are recognized, according to the ease with which a crystalline sublimate is obtained. Although the m. p. is of little significance for the sublimation process, the difference between the sublimation temp and the m. p. is

a deciding factor as to whether or not well-defined sublimate can be obtained. Tables are given showing for a no. of drugs the m. p., initial and optimum sublimation temps., and the nature of the sublimate, and these data, in connection with the crystal form of the sublimate, serve in most cases for identification. Sepn. of a mixt. of such compds. by sublimation is rarely possible, as for the majority the sublimation temp. lies between 100° and 140°.

B. C. A.

Crystallographic investigation of microsublimates of synthetic drugs. W. HAAS. *Mitteilungen, Zinnich Festkörper*, 1930, 63-119.—Details are given of the detn. of the optical properties of sublimate, and tabulated data are presented for a no. of synthetic drugs, comprising the appearance and crystalline form of the sublimate, the extinction, interference colors, optical character, dispersion and α , n , β and γ .

B. C. A.

Infusum radices petasanthae. K. H. BATTER AND K. HENCK. *Pharm. Zentralblatt* 71, 513-8 (1930).—Infusions of decoct. below 140°, prepd. by the method of D. A-B VI, contain the bulk of the alkaloids present as detd. in the original drug by the D. A-B VI method (ammoniacal ether extn.). More concd. infusions contain less of the original alkaloids, thus a 1/2 infusum contains only about one half. Decoct. is worse of dil. infusions results in considerable decrease of alkaloid content. Decoctions prepd. by the above method contain the same amts. of alkaloid as infusions of equal concn., and prolonged treatment does not raise the yield in either case. When the residue from a concd. infusion is again extd., a further quantity of alkaloid is removed, but the total yield remains deficient. Infusion of such a residue with dil. HCl brings the total yield almost to that indicated by the D. A-B VI detn. Infusion of the original drug with dil. HCl (2 g. with 40 cc. of water and 10 cc. of 0.5 N HCl, followed by a 2nd infusion with 50 cc. of water) gives, however, higher values for the alkaloid content than does the German official method, the figures from the latter representing only about 57% of the alkaloid present. The original drug therefore, contains alkaloid in a form (possibly glucosidic) that is removed neither by infusion nor by ammoniacal ether extn.

B. C. A.

Chemical composition of the [Russian] tobacco crops of 1927 and 1928. V. BALASHOVA POPKOVA AND I. ZAPOLSKII. U. S. S. R. State Inst. Tobacco Inv., *Bull.* 64, 82 pp (1930).—A large no. of data for tobaccos grown in different parts of Russia are given.

B. C. A.

Steam-air method of treating tobacco. P. K. DOKHOV, I. I. LOSEV, A. O. RYFFIK AND A. A. SHURUK. U. S. S. R. State Inst. Tobacco Inv., *Bull.* 64, 34 pp (1930).—By the method described, which consists in treating the leaves with steam and air, and effects uniform moistening and sorting, the time occupied is reduced to a few min., all the conveying of the tobacco being by mech. devices. At the same time the nicotine (recovered as sulfate), NH_3 , and MeOH are considerably diminished in amt., and the flavor is improved. The process has been tested on a semi ind. scale with satisfactory results.

B. C. A.

Polyphenols of tobacco. A. SHURUK. U. S. S. R. State Inst. Tobacco Inv., *Bull.* 69, 7-13 (1930). cf. C. A. 24, 2235.—From the ppt. formed by addn. of basic Pb acetate to the aq. ext. of unfermented tobacco (Tyk Kulak), a substance may be isolated which dissolves readily in water, exhibits tanning properties and gives an intense green coloration with FeCl_3 . It does not reduce Fehling soln., but yields a reducing methyl-pentose (rhamnose) on acid hydrolysis, which takes place only with difficulty. The substance is apparently a glucoside, another of its components being a decape, contg. both caffeic and quinic acids. Treatment with alkali hydrosulfide soln. yields protocatechuic acid.

B. C. A.

Inositol in tobacco. A. SHURUK. U. S. S. R. State Inst. Tobacco Inv., *Bull.* 69, 15-8 (1930). cf. preceding abstr.—The aq. ext. of fermented tobacco (Tyk Kulak) contains a considerable proportion of inositol and this may play a part in the biochem. processes resulting in the formation of nicotine and may serve to explain the presence in the nicotine mol. of a 6-membered ring which is absent from the protein mol. When tobacco seeds are treated with dil. mineral acid and the resulting soln. is pptd. with alkali, phytin is obtained to the extent of 7-10% on the dry matter of the seeds.

B. C. A.

Acids of tobacco. II. A. SHURUK AND PLATNICKI. U. S. S. R. State Inst. Tobacco Inv., *Bull.* 69, 19-26 (1930). cf. C. A. 23, 4771.—The ethereal ext. of fermented tobacco (Samsun) contains citric, caffeic and probably chlorogenic acids. Although the presence of quinic acid has been observed by various authors, it could not be detected and it may form one of the components of chlorogenic acid.

B. C. A.

Content of formic acid in tobacco. A. SHURUK AND S. KASHIRIN. U. S. S. R. State Inst. Tobacco Inv., *Bull.* 69, 29-32 (1930).—By using a modification of Fincke's method

for detg HCOOH (C. A. 7, 2983), it is found that this acid is present in both fermented and unfermented tobacco to the extent of 0.02-0.17%. Tobacco of high and low grades contain rather more than those of medium quality. Fermented tobacco are richer in the acid than unfermented ones, the proportion increasing with the energy of the fermentation process. Both HCOOH and its salts diminish the strength of flavor of tobacco. B. C. A.

Determination of the essential oils of tobacco. V. V. SEMENOVA. U. S. S. R. State Inst. Tobacco Inv., *Bull.* 69, 47-53 (1930).—The following procedure does not det. the total essential oils in tobacco, but serves for obtaining a comparative measure of the proportions present in different tobacco. Five g. of the finely-divided tobacco, heated in a flask in a boiling water bath, is distd in a current of steam under definite conditions, 4 cc. of water being distilled from a small boiler to furnish the steam. Air is drawn through the app. at const. rate (30 bulbles in 15 sec.) during the whole time of the distn. and for 70 min. afterward and passes through 3 washing vessels each contg 3 cc. of H_2SO_4 (d 1.84). The whole of the acid, together with washing water, is made up to 20 cc. and mixed, and is then compared colorimetrically with standard solns. of tobacco essential oils in H_2SO_4 . B. C. A.

Free nicotine or free alkalinity of tobacco. A. SHIMUK AND M. CHIMURA. U. S. S. R. State Inst. Tobacco Inv., *Bull.* 69, 65-9 (1930), cf. C. A. 24, 2235.—The content of nicotine in tobacco is related to the reaction of the material. Distn. in a current of steam of tobacco which has not been rendered alk. yields high proportions of nicotine (37-40%), which increase progressively as increasing amts. of alkali are added. 1 xtn. of tobacco with light petroleum gives considerably smaller yields of nicotine. The proportion of free nicotine present is much greater than that calcd. theoretically from the degree of hydrolysis of nicotine salts and the disocn. const. B. C. A.

Steam distillation of nicotine from tobacco. A. SHIMUK AND M. CHIMURA. U. S. S. R. State Inst. Tobacco Inv., *Bull.* 69, 71-80 (1930).—Addn. of CaO and NaCl to tobacco not only shortens the time of steam distn., but results in increased yields of nicotine. B. C. A.

Estimation of hexamethylenetetramine methylenecitrate and of hexamethylenetetramine in the presence of each other. D. MAROTTA AND F. DI STEFANO. *Ann. chim. applicata* 21, 150-4 (1931).—The method of estg. these 2 compds. in the presence of each other depends on the fact that the methylenecitrate is decomposed by alkali with liberation of CH_2O , while $(\text{CH}_2)_6\text{N}_4$ is stable. Dissolve 0.4-0.5 g. of the mixt. in 100 cc. H_2O , neutralize with 2 N NaOH , distil the liberated CH_2O into another flask, and det. according to Romijn (*Z. anal. chem.* 36, 18 (1897)). To the remaining liquid add 100 cc. water and 50 cc. of 5% H_2SO_4 , collect the CH_2O liberated from the base and det. as above. A. W. COFFEE.

The preparation and properties of colloidal aluminum hydroxide for medicinal use. J. J. L. ZWICKER. *Pharm. Weekblad* 68, 328-43 (1931).—A series of $\text{Al}(\text{OH})_3$ prepn. was made for the purpose of obtaining a product with the properties of "Aluocol," a com. $\text{Al}(\text{OH})_3$ prepn. used in the treatment of gastric disturbances. The desirable properties which distinguish "Aluocol" from other $\text{Al}(\text{OH})_3$ prepn. are ready dispersion in dil. HCl , slow neutralization, and gelatinization with concd. HCl . Such a prepn. was eventually obtained by the method of Kohlschütter (C. A. 14, 242, 17, 2943) in which a solid cryst. Al salt is treated with NH_4OH . A product practically identical with "Aluocol" was obtained as follows: Powd. $\text{Al}_2(\text{SO}_4)_3$ was allowed to stand 4-5 days at room temp. or 18 hrs. at 40° with 10 vols. of 5% NH_4OH contg. 5% $(\text{NH}_4)_2\text{SO}_4$, the ppt. washed with tap water and finally with distd. water, then dried at 40° and finely powdered. A. W. DOX.

Micro melting-point determinations of drug products. L. KOFLER AND W. DERNBACH. *Arch. Pharm.* 269, 104-14 (1931).—A report of results obtained in the examn. of some 54 different D. A. B. VI drug products by the use of Kofler and Hillick's app. (C. A. 25, 1121). Except with dionine, scopolamine-HBr, terpine hydrate and cocaine- HNO_3 , the m. pts. observed by the micro method showed satisfactory agreement with those obtained with the usual macro procedure. The exceptions to the rule find probable explanation in the water of crystn. or constitution of the substances noted. W. O. E.

Essential oil of Taiwan cedar. KINZO KAFUKU AND RYO KATO. *Bull. Chem. Soc. Japan* 6, 65-74, *Chem. News* 142, 273-6 (1931).—The wooden chips of *Taiwania cryptomerioides* Hayata give on steam distn. 0.23% of an oil, consisting chiefly of sesquiterpenes and sesquiterpene alc. The sesquiterpenes in the oil are mainly cadinene, with a little humulene and caryophyllene. From the cadinene a new oxidation product, $\text{C}_{15}\text{H}_{22}\text{O}_2$, m. 142-3°, was obtained. The sesquiterpene alc. in the oil, for which the

authors suggest the name "taiwanol," closely resembles cadinol of Semmler and Jonas. From cadinol no cryst. phenylurethan could be obtained, but with taiwanol it affords a cryst. phenylurethan, m. 134-5°, and furthermore, it gives a neutral oxidation product m. 170° if oxidized with KMnO_4 in AcMe soln.

Color of tinctura valerianae aetherea. G. BERNING *Apoth. Ztg.* 46, 47-8 (1931).—The tincture of valerian freshly prep'd is bright yellow, and, as the D. A. B. VI states, becomes darker with age.

Reaction of cod-liver oil and castor oil. LAD. ECKERT *Pharm. Zentralhalle* 72, 209-12 (1931).—The tests specified in the U. S., Italian, German, Dutch and Hungarian Pharmacopoeias for cod liver oil are reviewed. The behavior of this oil, as also that of castor oil when treated with aldehydes and coned H_2SO_4 , is described. Thus, if 1 drop of cod liver oil is dissolved in 1 cc. strong FeOH , then treated with 3 drops of 1% aldehyde, and the mixt. thereupon underlaid with 1 cc. coned H_2SO_4 , a color zone develops. After shaking the liquid is exam'd under the quartz light. The various color changes are described when different aldehydes are used.

Biological differentiation of morphine-containing tinctures of varying potency. K. FEIST AND F. KLATT *Apoth. Ztg.* 46, 30 (1931).—1 vap. 15 g. of the sample (opium tincture) to dryness on the steam bath, add 3 cc. dil. AcOH and again evap. to dryness to expel excess acid, then dissolve the residue in 15 cc. dist'd H_2O . Inject the clear soln. subcutaneously in the back of a white mouse, and note the typical tail reaction and subsequent death of the animal. Addnl. samples are prep'd as above, from which 1 cc. and 0.5 cc., resp., are injected in other mice. If the 1st died in 3 min., the 2nd in 7 and the 3rd in 22 min., a basis of comparison is established for opium tincture, which is about 10 times stronger in morphine content than cholera drops, the 2 preps. under consideration in the present paper.

Tincture diaculation. H. DREBIN *Pharm. Ztg.* 76, 400-2 (1931).—The results obtained in the prep'n of a few special tinctures, notably of sabadilla, by the process of diaculation are reported with favorable comment.

The reactions of oleum jecoris and oleum ricini. L. ECKERT, *Magyar Gyógyszer-tud. Társaság Ertetése* 7, 115-20 (1931).—One cc. oil is dissolved in 1 cc. coned alc., 3 drops of 1% alc. solns. of the aldehyde are added and 1 cc. coned H_2SO_4 is poured in a layer underneath. The colors of the rings and liquids given by oleum jecoris with the different aldehydes before and after thorough mixing and in the light from a quartz lamp (Q) are, resp. (1) HCHO , rust brown ring, brown and (Q) light yellow liquid, (2) PACH_2CHO , reddish brown ring, rust red and (Q) light pink liquid, (3) BzH , rust red ring, rust red and (Q) reddish yellow liquid, (4) 2 furaldehyde, ring dark red underneath and dark green on top, liquid violet red and (Q) Bordeaux red, (5) α -unsaturated aldehyde, dark brown ring, dark violet and (Q) Bordeaux red liquid, (6) α -unsaturated aldehyde, dark red ring, blood red and (Q) reddish violet liquid, (7) vanillin, ruby red ring, carmine-red and (Q) reddish violet liquid, (8) cinnamaldehyde, dark red ring, then dark reddish violet and (Q) brownish red liquid, (9) piperonal, ruby red ring, carmine red, and (Q) orange-to brick red liquid. The colors given by oleum ricini under the same conditions were (1) (2) or (3) brown ring, reddish brown and (Q) light olive green liquid, (4) dark brown and dark green ring, dark red and (Q) liver-red liquid, (5) carmine red ring, dark violet and (Q) liver red liquid (6) yellow and red ring, blood red and (Q) grayish olive-green liquid, (7) ruby red ring, carmine-red, and (Q) liver red liquid (8) dark red ring, carmine red and (Q) olive-brown liquid; (9) ruby red ring, carmine red and (Q) brownish yellow liquid.

Proposed standards for fennel fruit. DR. GRAAF *Heil u. Genußpflanzen* 12, 41 (1929), *Quart. J. Pharm. and Pharmacol.* 3, 252.—Recommendations are put forward for the fixing of an international standard for fennel fruit. The chem. requirements with which the fruit should comply are as follows: Bitter fennel, first quality—volatile oil at least 5%, ash not more than 8%; sand content not more than 0.2%. Sweet fennel, first quality—volatile oil at least 2.5%, ash and sand content as for bitter fennel.

Syrup of ferrous iodide. G. KULLERUD *Norg. Apot. Tidss.* 29 (1929); *Quart. J. Pharm. and Pharmacol.* 3, 544.—Small quantities of syrup kept in half filled bottles in the dark became pale yellow after 2 months. On exposing these bottles to bright sunlight the color was not affected, but a portion of this syrup in a completely filled bottle became colorless after 3 days' exposure to sunlight. The author concludes that, in completely filled vessels, syrup of FeI , does not become colored, in opened bottles the syrup can be kept for nearly 2 months. Decompos. is caused by the O of the air, and not by light.

Determination of chlorine numbers of drugs. J. STAMM *Farm. Notisblad* 39,

21; *Quart J Pharm and Pharmacol*, 3, 253-4.—To det. the Cl no ext 0.25 g of finely powd drug with boiling water and make up to 250 cc. Transfer 3 portions of 5, 10 and 15 cc, resp., to stoppered conical flasks and make up to 25 cc with water. To each flask add 25 cc. of a filtered 0.1 N soln of chlorinated lime, followed by 2 cc. of 20% H_2SO_4 . Let the flasks stand out of direct sunlight for 5 min, add KI, and after a further 5 min titrate with thiosulfate soln. Make a blank detn at the same time. The Cl no is the no of cc of thiosulfate used up by 0.025 g of the drug. The chlorinated lime soln must be used fresh but may be kept in the form of a more concd filtered soln contg about 20 g of active Cl per l. G G

Chlorine numbers of drugs. K KILLINEY. *Farm. Notisblad* 39, 125(1930), *Quart J Pharm and Pharmacol* 3, 254.—K records the Cl nos of a large no of drugs, detd according to the method of Stamm (see preceding abstract). The Cl no depends upon the age, ripeness and manner of preservation of the drug. In many cases there appeared to be a difference between fresh and old samples, though in general the no of samples examd was not sufficient to draw final conclusions in this regard. G G

Ethylene as a dental anesthetic. HARRY M SELDEN. *Dental Cosmos* 73, 454-7 (1931).— C_2H_4 has its place as an anesthetic in major surgery in hospitals where adequate precautions can be taken against explosions. C_2H_4 in mixts with other gases (CO_2 , O_2) is not suitable as a dental anesthetic because of the danger of either explosion or asphyxia. JOSEPH S HILBERY

The principles and methods of biological assay. J M WATT. *J S African Chem Inst* 13, No 2, 41-7(1930). E H

Recent advances in medicinal preparations in Germany. KONRAD SCHULZE. *Mfg Chemist* 2, 103 5(1931).—Brief discussions are given of *Eupaverine* (1-(3,4-methylene-dioxybenzyl)-3-methyl-6,7-methylenedioxyquinoline), *Fantan* (phenylethanolonyl urethane, $\text{PhC}_6\text{H}_4\text{NCONHCO}_2\text{Et}$), *Iodine sticks* (a colloidal I prepn), *Gullajod* (an I-contg tablet), *Lacarnol* (a "heart-hormone" prepn consisting of specific parts of the heart muscle together with skeletal muscle and other body tissues), *Pantocaine* (*p*-butylaminobenzylidimethylaminoethanol), and *Sympathol* (*p*-hydroxyphenylethanol-methylamine). E H

Posterior pituitary preparation. C G MACARTHUR. *Science* 73, 448(1931).—Beef lobes were repeatedly treated with acetone, ether, petroleum ether and a mixt of 70 parts alc, 25 AcOH and 5 water. The final ext was moderately sol in boiled distd water, easily sol in dil acids and alkalis and was pptd by Cu and Zn salts. It seemed to be a polypeptide containing fabile S. It gave a strong reaction on blood vessels and isolated uterus. MARY I. LEAR

Limitations of phenol-coefficient tests in determining germicidal activities. GRO F. LEONARD. *J Infectious Diseases* 48, 358-66(1931).—The phenol coeff. test is limited in value when applied to substances, like certain Ag preps, that have a high bacteriostatic action. By adding the transfer test to the regular phenol coeff test, one can readily det which products are actually germicidal and which have only a restraining effect on bacterial growth. A method for testing growth restraining action is described which depends on color changes rather than turbidity to det growth and is therefore applicable to substances that cloud the media. Both the mild and strong Ag preps have high bacteriostatic action, restraining typhoid and staphylococcus cultures in dilns of 1:5,000 to 1:20,000. Colloidal Ag preps have a feeble bacteriostatic action, restraining growth only in dilns of 1:50 or 1:100. JULIAN H LEWIS

Organic antiseptics. A bacteriologic study of a new series. W. ERIC DRAKE and FITZGERALD DUNNING. *J Infectious Diseases* 48, 366-71(1931).—A series of derivs of resorcinolsulfonephthalein and of halogenated resorcinolsulfonephthalein was prepd and tested bacteriologically. The introduction of halogen atoms did not greatly increase the bactericidal activity of the parent compd, resorcinolsulfonephthalein. The Hlg deriv is considerably more bactericidal than the parent substance. The increase in bactericidal activity is greatest when both Hlg and halogen atoms are introduced. The various halogens when added along with Hlg increase the bactericidal activity in the order, Cl, Br, I. Di Hlg derivs are not greatly more bactericidal, but are considerably more toxic than the corresponding mono- Hlg compds. The Na salt of monohydroxymercuroiodoresorcinolsulfonephthalein proved to be by far the best compd of this series from a therapeutic standpoint. JULIAN H LEWIS

The influence of irradiation on adrenaline. I. The action of ultra-violet rays (of artificial light) and of sunlight on adrenaline. BUNJIRO TEBATA and RYOSICHI ITO. *Folia Pharmacol Japon* 12, No 1, 74-81, Breviaria 5-6(1931).—The action of adrenaline decreases with an increase in color of its soln. When adrenaline is treated with ultra-violet light, there is an increase in potency. The soln also gains strength

when irradiated with artificial light, but when 10 sunlight units are applied the action is decreased. The action of the soln does not correspond to the color developed. Exposure to sunlight increases the adrenaline action, but over-exposure is injurious. The color developed is no indication of the strength. Adrenaline soln left unstoppered loses its strength, but this may be revived by irradiation with Röntgen, ultra violet or sunlight rays. G H W LUCAS

Studies on "ovarmina" (Erbe). V. CLANCE *Boll. soc. sci. nat. sper.* 6, 57-9 (1931). — "Ovarmina" when tested by the Allen Dossy technique contains not less than 150 rat units per cc. PETER MASUCCI

Cod-liver oil of Japanese Dorabees "Tara." X. K. KAWAI, *J. Pharm. Soc. Japan* 51, 177-183 (1931), cf. *C. A.* 25, 381. F I NAKAMURA

Pharmacognostic study of *Aralia chinensis* L. var. *glabrescens* Matsum. T. MURVADA, *J. Pharm. Soc. Japan* 51, 200-6 (1931). F I NAKAMURA

The iconoclast. FREDERIC WALKER, *J. Chem. Education* 8, 835-85 (1931). — W. calls attention to the lives and works of Paracelsus and Galen showing how each took exception to the established authority of medicine of his day and pleaded that progress is possible only as the result of expt. E R SCHERZ

Behavior of quinine salts with some acids and salts. V. MACCHI *Boll. chim. farm.* 69, 497-500, 503-4, 507-10, 529-40, 543-6 (1930). — A detailed description of the reactions of quinine salts with some org. and inorg. acids and salts is given. G S

Practical contribution to the preparation of solutions for hypodermic use. GIRO BARATTEN *Boll. chim. farm.* 69, 1029-31 (1930). — The app. permits the sterile filtration, measuring, and filling in vials of solns for hypodermic injection. The filtration can be carried out with suction and the procedure can be performed in an atm. of an inert gas. G SCHWOCH

Somnival Erba and Passiflorine Béolet. VITTORIO ZANOTTI *Boll. chim. farm.* 69, 1031-2 (1930). — Somnival Erba and Passiflorine Béolet are fluidexts of *Passiflora incarnata*, which contain an essential oil and tannin. The reactions given by the essential oil are described. G SCHWOCH

The Italian Pharmacopoeia in its fifth edition. SERAFINO DEZANI *Giorn. farm. chim.* 80, 101-11 (1931). G SCHWOCH

Chemical constituents of the bark of *Melicope erythrococca*. T. G. H. JONES AND M. WHITE *Proc. Roy. Soc. Queensland* 41, 154-7 (1929). *Australian Sci. Abstracts* 9, 28. — The results confirm the view of Cohen (*C. A.* 2, 1447) that lupeol is a mixt. of 2 substances, one, however, apparently present to a much greater extent than the other. J D S.

Essential oils from the Queensland flora. I. *Baetkea virgata*. T. G. H. JONES AND M. WHITE *Proc. Roy. Soc. Queensland* 42, 49-51 (1930). *Australian Sci. Abstracts* 9, 40. — The essential oil was from one lot of leaves collected near Gindina; yield 1% d₄²⁰ 0.9021, $[\alpha]_D^{20} +18^\circ$, $n_D^{20} 1.4742$. The principal constituents were d- α -pinene, cineol (30%), aromadendrene, a sesquiterpene alc. and possibly pinocaradiol. J D S.

The preparation of the tinctures of plants. B. CRISTINI *Ann. soc. pharm. chim. Sao Paulo* 2, 81-8 (1931). — A critic of methods in use with recommendation that the procedure of the I. S. P. be followed. M. H. SOULE

Preparation used in the treatment of addiction to morphine, heroine and cocaine. F. MASTRANGIOLI *Ann. soc. pharm. chim. Sao Paulo* 2, 106-11 (1931). — A discussion with an attempt to det. the compn. of these substances. M. H. SOULE

The handling of balsams of Peru. F. PITOMBO *Ann. soc. pharm. chim. Sao Paulo* 2, 138-43 (1931). — The method of collecting pure balsam of Peru and the phys. const. of this substance are given. The materials frequently used as adulterants are discussed. M. H. SOULE

New method of evaluation of cantharides. I. J. OLIVELLA RIV. *Anales soc. españ. f. quim.* 29, 164-9 (1931). — The dry material is extd. to completion with a mixt. of acetone and HCl, the soln. evapd., and the crystals obtained are purified by 90° alc. and petroleum benzene satd. with HCl and cantharidin, then recrystd. from acetone and washed with petroleum benzene. Material from all parts of the world varied in cantharidin content from 0.000 to 1.825%, differing markedly from the value 0.60% accepted by the Brussels International Conference. The cantharidin obtained m. 216°. E. M. SYMES

Australian sandalwood oil. W. H. SNODGRASS, *Chemist and Druggist* 114, 89 (1931) cf. *C. A.* 25, 773. — The statements of Walsmsley are refuted, those of Penfold, Marr, etc., are upheld. S. WALDBOTT

Dog fat. J. PRITZKER AND ROBERT JUNGKENS *Pharm. Acta Helv.* 6, 55-63

(1931)—Dog fat has long been used in Swiss and Austrian domestic medicine. In 7 samples of supervised origin and 2 com. samples, constants were detd and the results tabulated. They largely agree with results from hog fat except that the Bömer no. (m-p difference, cf *C. A.* 8, 1174, 2133) is lower in dog fat (67-69) than in hog fat (not below 71). An admixt. of hog fat to dog fat may thus be detd. to a certain extent. Adulteration of 1 sample of dog fat with hog fat was established on the basis of the Bömer no. as well as by the method of Grossfeld and Sammer (*C. A.* 24, 5174) for the sepn. of the solid from the liquid fatty acids, and by means of the method of Kreis and Roth (*C. A.* 7, 1815) of fractional crystn. The Grossfeld method also demonstrated for dog fat (as previously shown for hog fat) the absence of unsatd. solid fatty acid. S. WALDBOTT

Comments and criticisms on the (British) pharmacopoeia revision. A. J. JONES *Pharm. J.* 125, 647 (1930), 126, 10 (1931), cf *C. A.* 25, 1919.—In the prepn. of *infusum digitalis*, the use of a coned infusion should be expressly prohibited. Similarly, in the use of industrial methylated spirits for the prepn. of galenicals the cases where its use is objectionable should be specifically noted. Addnl. suggestions are made: *Syrupus ferri phosphatis compositus*. E. R. MEACHEM *Ibid.* 126, 10, 31.—To prevent pptn. when the formula given by the subcommittee is followed, the partial substitution of H_2PO_4 for Li_2PO_4 has been suggested (cf J. H. Franklin) but was not adopted. New expts. show that sirups contg. 2% + of dil. H_2PO_4 , Brit. Pharm. Codex, when exposed to bright light for 2.5 months remained clear, the others not contg. H_2PO_4 gave a heavy ppt. Use the formula of the subcommittee, modified: Fe (wire) 4.3 g., H_2PO_4 (d. 1.75) 28 cc., dil. H_2PO_4 20.0 cc., $CaCO_3$ 8.9 g., $KHCO_3$ 1.0 g., Na_2HPO_4 1.0 g., cochineal 3.4 g., sugar 700 g., H_2O to make 1000 cc. Add the Li_2PO_4 to the mixed solns. just before filtration into the cold sirup. *Cataplasma kaolin.* J. H. RIDGEWAY *Ibid.* 126, 31.—To avoid retention of the total H_2O of reaction in the subcommittee's method of mixing $B(OH)_3$ and glycerol, dissolve the $B(OH)_3$ in the glycerol and heat with vigorous stirring to expel as much H_2O as possible before adding the kaolin. Estimation methods. J. F. LIVERSFEGE *Ibid.* 126, 31.—Suggestions are made on the definition and method of detn. of the sapon. no. in pharmacy, on "normal pressure" reading in measuring NO gas from sweet spirit of nitre, etc. In the detn. of As , the use of strips of $HgCl_2$ paper in place of a cap permits of quant. comparisons when allowed to act overnight in the cold. The use of a hot plate is condemned, as it gives erratic results. Practical notes. J. J. BLACKIE *Ibid.* 126, 4, 7, 16-18, *Chemist and Druggist* 114, 20-31 (1931).—Detailed comments and suggestions are made and discussed on exts., glycerins, infusions, ointments, tinctures and waters. Pharmacopoeia revision. DAVID HARLEY *Pharm. J.* 126, 109-12, *Chemist and Druggist* 114, 185 (1931).—An extensive discussion by H. and others. The substitution of other oils for olive oil in *lunientum emphorae* proposed by the subcommittee is strongly deprecated. S. WALDBOTT

Liquor plumbi subacetatis fortis. J. E. WOODHEAD *Pharm. J.* 126, 28 (1931).—Examn. of 38 samples of semi-official lotions contg. definite vols. of Goulard's ext., Brit. Pharm. (A) indicated excess of Pb present. In 4 com. samples of A, the required min. of Pb subacetate was exceeded by an av. of 16.0%. W. suggests that if the present formula for A be retained, d. should be between 1.275 and 1.290, and the Pb subacetate content between 26.5 and 28.0%, which figures conform to actual conditions. S. WALDBOTT

Psyllium seeds. HUBERT B. Figg *Pharm. J.* 126, 29 (1931).—The seeds of *Plantago psyllium* L., fleawort, are used as a mild natural laxative in chronic constipation. The mucilage test differentiates the (worthless) pale variety of seeds from the medicinal dark variety. Cover a small quantity of the seeds in a test tube with about 3 or 4 times their wt. of H_2O . After a short time only the dark seeds have swelled to a semisolid mass which will remain in place when the tube is inverted. S. WALDBOTT

Phenol-water mixtures. G. J. W. FERRY *Pharm. J.* 126, 58 (1931).—Detns. of f. p. of $PhOH-H_2O$ mixts. (78.1-96.17% $PhOH$) gave results which agreed well with those of Rhodes and Markley (*C. A.* 16, 670). The f. p. of the liquid carbohc acid (80% $PhOH$, 20% H_2O) proposed by the subcommittee for the new Brit. Pharm. is 3.1°; the sp. gr. close to 1.0615. F. suggests that the limiting values for H_2O 19.4-21.9%, sp. gr. 1.062-1.060, f. p. 2.3-3.4° be adopted for the new liquefied carbohc acid. S. WALDBOTT

The preservation of sirupus acaciae, Brit. Pharm. Codex. H. A. TURNER *Pharm. J.* 126, 160 (1931).—With mucilage of acacia and sirup of acacia fermentation takes place quickly when preservatives are not added, but addn. of $BzONa$ or $CHCl_3$ prevented fermentation, but not formation of acidity. As the use of $BzONa$ may be

objected to, the use of 0.25% CHCl_3 is recommended (cf. *C. A.* 23, 5010; 24, 202, 5038) S WALDBOTT

The pharmacist and the pharmacopeia. W. A. KNIGHT. *Pharm J.* 126, 161 (1931) — In some cases the pharmacopoeial tests are unnecessarily strict and too elaborate for the ordinary pharmacy, e. g., the As test (5 p. p. m.) in NaNO_2 of "reagent" quality, where a simpler darkening test with soln. of H_2S or $(\text{NH}_4)_2\text{S}$ should suffice. K. suggests that the Brit. Pharm. tests be of 2 classes: (1) Delicate and complicated but important tests, e. g., those for biol. products which are bought in sealed containers under guarantee. (2) Tests of identity and purity which are within the pharmacist's scope to perform S WALDBOTT

Linimentum camphorae, modified. A. N. WINTER. *Pharm J.* 126, 184 (1931) — Rub down camphor, 1 oz., with essential oil of camphor 0.5 oz. and add sufficient olive oil to make 10 oz. A clear nonpptg. and satisfactory liniment results. The formula is recommended for adoption in the Brit. Pharm. S WALDBOTT

Notes on some leguminous seeds. G. I. TARRER. *Pharm J.* 126, 211-2 (1931) — Pharmacognostical descriptions and photographs are given of the seeds of *Cassia fistula*, *Tamarindus indica* and of *Indian* and *Alexandrian sesna* S WALDBOTT

Extracts obtained in the manufacture of ethereal oils from conifers (Donaov) 22. New and improved methods in the lormation of pharmacologically important amines II. Synthesis of β -arylethylamines from aromatic aldehydes and carboxylic acids (KINDLER) 10. Problems of the technical manufacture of papaverine [and laudanum] (KONER) 10. Constitution of hydroxyoleic acid from oil of ergot (MATTHEWS, KOSCHNEA) 10. The properties and effects of a new class of organic Br compounds (HILMANN, FREUND) 11H. Amino alcohols (HARTUNG, MUNCH) 11H. Condensation products [for use in pharmacy] (Fr. pat. 698,525) 25.

GRAPE, VICTOR. *Narkotische Genuasmittel, Drogen, Gewürze, Harze*. Stuttgart. C. E. Poeschel Verlag. 688 pp. M 44.50

HALE, WALTER, WM. *Materia Medica, Pharmacy, Pharmacology and Therapeutics*. 20th ed., revised by A. H. Douthwaite. London. Churchill. 712 pp. 10s. 6d. net

HOUSE, A. G. *Bacteriological Control of Milk: a Practical Guide for Media Preparation and Milk Testing*. Reading. The Natl. Inst. for Research in Dairying. 59 pp. 3s. 6d.

Epitome of the Pharmacopeia of the United States and the National Formulary. Prepared under Authorization of Council on Pharmacy and Chemistry of the Am. Med. Assoc. 4th ed. Edited by R. A. HATCHER, FRANK E. IRONS, DONALD SOLL-MANN and WM. A. PUCKNER. Chicago. Am. Med. Assoc. 238 pp.

TSCHIRCH, A. *Handbuch der Pharmakognosie*. 2nd ed., enlarged. Lfg. 6. Leipzig. Bernhard Tauchnitz. M 8. Cf. *C. A.* 25, 1315

Medicinal composition containing agar. FERDINAND W. NITARDY (to E. R. Squibb and Sons). U. S. 1,709,804, April 7. In prepg. liquid products contg. agar in a quantity greater than that which would normally form a firm jelly in the presence of water, the agar is dissolved in hot water in the proportions required to form a firm jelly, and the soln. is cooled and stirred to produce a semi-liquid mass of comminuted jelly particles and the latter are emulsified with a mixt. of a laxative oil and water.

Irradiating medicaments, foodstuffs, etc. OSKAR RIED. Austrian 121,657, Oct. 15, 1930. The activating effect of ultra violet rays on medicaments, foodstuffs, etc., is enhanced by subjecting the materials to the action of other rays, e. g., Röntgen rays or rays of long wave length, before, during or after irradiation with ultra-violet rays.

Apparatus for making physiological salt solution in large quantities. RUDOLF A. HARTMANN. Ger. 522,495, Aug. 3, 1927.

Calcium carbohydrate compounds. ERNST STERN. Ger. 522,302, Oct. 13, 1925. Assimilable products of therapeutic value, mostly sol. in water, are prepd. by treating solns. or suspensions of carbohydrates of the sugar type with $\text{Ca}(\text{OH})_2$ at a temp. not exceeding 60°. The concn. of the $\text{Ca}(\text{OH})_2$ is gradually increased during the initial stage of the process, e. g., up to 5% of CaO , calcd. on the carbohydrate. After about 12 hrs., the product is dried *in vacuo* at a temp. not above 70°. The feeble alk. of the product may be removed by incorporating a sol. acid salt, e. g., NaH_2PO_4 .

Cyanamide-formaldehyde condensation product. I. G. LARSENING A.-G. (Hans Schmidt, inventor). Ger. 522,656, Sept. 11, 1924. Cyanamide is condensed with

CH_2O directly or in aq soln, yielding a nitrogenous product useful as an intermediate for drugs, etc. The cyanamide may be prepd from an aq soln of its Ca compd, and the condensation may be accelerated by means of alkali. Examples are given. Cf C. A. 25, 1645

Ephedrine and structurally similar substances. RICHARD H. F. MANSKE (to Eli Lilly & Co.) U. S. 1,799,110, March 31. Numerous details and examples are given of the prepn of ephedrine and structurally similar substances by a process which consists in: (1) the condensation of a diketone of the general formula RCOCOR' with a primary amine of the general formula $\text{H}_2\text{NR}'$ in which R is a benzenoid radical, R' is $\text{H}(\text{CH}_2)_x$, with x any desired integer including zero, and R' is an org radical which contains a C atom by which it is linked to the N atom, (2) the subsequent hydrogenation of the product of the first step. Ten detailed examples of the process are described.

Vitamins. TIBIRIUS REITER Jr 697,367, Jan 30, 1930. Substances contg ergosterol are vitaminized by treatment with ultra-violet light, only rays the wave length of which is above 2,800 Å are used.

Gland extracts. GEORGES L. A. PILLERIN Fr 698,564, June 26, 1930. In extg opotherapeutic products from animal glands, the glands are stabilized at the moment of killing so that the products are preserved in full vigor. Thus the glands are removed when the animal is killed, kept for 10–20 min at 40° in the absence of air, then cooled to about 4° at which temp stabilization is definitely established.

Anesthetics. SCHIRING-KAULBAUM A-G (Walter Schoeller and Hans G. Allardt, inventors) Ger 522,064, Aug 11, 1928. Alkamine esters of aminobenzoic acid having an alkoxy group substituted in the nucleus, are prepd by the customary processes from monoalkoxy-nitro- or -amino-benzoic acids or their derivs. Thus, 4-nitro-3-methoxybenzoyl chloride may be heated in CaH_2 soln with $\text{Et}_3\text{NC}_2\text{H}_5\text{OH}$, and the product reduced with SnCl_2 and HCl , yielding diethylaminoethyl 4-amino-3-methoxybenzoate-HCl, m 150°. Examples are given also of the prepn of diethylaminoethyl 4-ethoxy-3-aminobenzoate-HCl, m 162–3°, and diethylaminoethyl 3-amino-4-methoxybenzoate-HCl m 223°.

Bactericides. THEODOR SADALITSCHKA and NÄHRMITTELFABRIK JULIUS PENNER A-G Ger 522,005, April 17, 1927. Use is made of water-sol products obtained by fusing esters of carbocyclic acids with sugar. A suitable product is prepd by fusing Me p-hydroxybenzoate 1 with sugar 9 parts.

Compositions of silicates with other compounds. HENRI BALAND Fr 698,543, June 18, 1930. Solid compns are made from alkali silicates sol in cold water and products such as antiseptics or dyes.

Face powder. YUZI NAGAI and TAKUZI NISHIMURA Japan 89,531, Dec. 11, 1930. The powder is a mixt of TiO_2 , talc, fibron prepd from HCl soln of silk, glycerol, H_2O and perfume.

Hair shampoo. HENRI CHAUDO Fr, 697,993, Aug 5, 1929. A hair shampoo contains CCl_4 , EtOH , acetone, PrOH and essential oils.

Fumigating rooms, etc. FARBWALZ GRS M B H Ger 522,316, Feb 19, 1928. In fumigating rooms, etc., by strewing them with cyanides unstable in the air, the evolution of H_2S from sulfide impurities in the cyanides is hindered by adding to the cyanides an oxidizing agent, e g, PbO_2 or MnO_2 . The action of the oxidizing agent may be enhanced by adding also a powd dehydrating agent, e g, CaCl_2 .

18—ACIDS, ALKALIES, SALTS AND SUNDRIES

E. M. SYMMES

Phosphoric acid. W. C. WELLER J. S. African Chem Inst 12, 50–7 (1929).—A review of the methods of manuf. E. H.

Concentration of phosphoric acid by submerged combustion. C. FEATHERSTONE HAMMOND Trans. Inst Chem Eng, Adv copy, March 24, 1931, 6 pp.—In the Hammond unit the burner is central in a refractory-lined combustion chamber, wholly submerged below the level of the liquid. The fuel gas and air are adjusted for complete combustion. Seger cones showed a temp. in the combustion zone of 1740°. The burner is made of non-corrosive metal cut from the solid, instead of from sheet or rod. Trapping the exhaust showed 0.02% loss of P_2O_5 in concg from 43° to 111° Tw. The acid is practically free from contamination, and the burner has a long life. High thermal efficiencies are obtained.

E. M. SYMMES

Improvements effected in the production of nitric acid by oxidation of ammonia.

GIACOMO FAUSER. *Chimie & industrie* 25, 556-69(1931)—See C. A. 25, 382.

A. PAPINEAU-COUTURE

New applications of ammonia. M. H. MERRILL. *Compressed Gas Manufacturers' Assoc. 18th Ann. Rept.* 1930, 28-37(1931)—At present about 32.5% of the synthetic NH_3 of the U. S. A. is employed in agriculture. For other purposes the percentages are: 17 for explosives, 12 for refrigeration, 20 for miscellaneous purposes and 51 for chemicals and acids. NH_3 is used as an alkali, a reducing agent, a source of HNO_3 , H and N, a fertilizer material and a means of prepping foods, treating rubber and purifying drinking water. A good bibliography is included. W. H. BOYNTON

Data and experiences in the manufacture of artificial nitrate. T. SZAFKA. *Technika* 11, 126-35(1930). cf. C. A. 25, 2525—Expts. were carried out on concn. 29.7° to 31.8° Bé HNO_3 and a waste acid contg. 69-70% H_2SO_4 , 5-10% HNO_3 and 20-25% water. The usual heating with steam dilutes the acid, therefore heating with warm air was tried successfully. The detailed literature of patents on the concn. of H_2SO_4 is listed. The boiler method is advised for artificial nitrate factories. The manufacture of NaNO_3 and its decomposition to HNO_3 are described. S. S. DE FINALLY

Some properties of commercial nitrates of lime. CAMILLE MATIGNON. *Compt. rend.* 192, 777-80(1931)—The 2 principal forms of com. nitrate of lime used in the fertilizer industry are the Norwegian, contg. 13% N and 22 to 23% H_2O ; and the ammoniacal nitrate of lime, contg. about 5% NH_4NO_3 and 14% water. Both varieties are very hygroscopic. A sample of the ammoniacal form increased in wt. 42% after 2-hr. exposure to the air, and 72% after 6 hrs., then dissolving. The possibility of spontaneous combustion of these products during shipment or storage was investigated. Thermochem. data are presented to show that decomposition of both forms is endothermic and cannot occur spontaneously at ordinary temp. Ammoniacal nitrate of lime, which is the more active form, was mixed with org. combustibles such as might be derived from the containers in which it is usually shipped, and the mixts. were heated gradually. Up to 300° the combustibles in the mixts. were unchanged. Combustion first occurred between 320° and 350°. Com. nitrates of lime, therefore, should not be susceptible to spontaneous combustion during shipping and storage. Actual absence of such accidents verifies this conclusion. There is no warrant for increased insurance rates based on spontaneous combustibility. R. H. LOMBARO

New process for the recovery of sodium nitrate in Chile. MAX SELIGMANN. *Z. Ver. deut. Ing.* 75, 513-6(1931)—Descriptive. E. H.

Separation of soda ash from brines of the Lake Tangar. S. Z. MAKAROV. *Zhur. Prikladnoi Khim.* 3, 1031-40(1930); cf. C. A. 24, 5438—Possibility of sepp. soda ash from Na_2SO_4 is shown. V. KALICHEVSKY

New procedure for manufacturing permanganates. LASLÓ DE PUTNOKY AND BÉLA DE BÖNÉSY. *Matematik és Természettudományi Értesítő* 47, 751-67 (in German 768-9) (1930)—Pyrolusites with a low SiO_2 content are dissolved in H_2SO_4 , forming MnSO_4 ; more ore and H_2F are added to the soln., producing MnF_2 , the electrolytic oxidation of which requires less current. As the electrolyte contains 10% H_2SO_4 and 5% H_2F , special diaphragms should be made, which are resistant to this acid mixt. Expts. proved that material contg. 90 parts fluorite and 5 parts cryolite formed into diaphragms and burnt at 750° is resistant, its loss of wt. amounting to only 0.1% daily. The anodic liquid should be worked up instantly, since it decomps. quickly. Permanganates of alk. earth metals can be made directly, those of alkali metals indirectly only. The process is therefore not economical for producing alkali permanganates. S. S. DE F.

Working up cyanide solutions to give potassium ferrocyanide and the preparation of potassium ferricyanide by treatment with compressed air and subsequent electrolytic oxidation. W. GLUUD, W. KLEMPF, K. KELLER, F. BRODKORF AND CL. DIECKMANN. *Ber. ges. Kohlenlech.* 3, 385-94(1931)—Cyanide solns. are converted into $\text{Ca}(\text{CN})_2$, which is converted into $\text{Ca}_2\text{Fe}(\text{CN})_6$ by treatment with FeSO_4 . From this, $\text{K}_4\text{Fe}(\text{CN})_6$ is obtained by double decomposition with K_2CO_3 . To obtain the ferricyanide, the method found most successful was oxidation of $\text{Ca}_2\text{Fe}(\text{CN})_6$ with compressed air in an autoclave, CO_2 likewise being present to remove $\text{Ca}(\text{OH})_2$, since the action is reversible in presence of an alkali. From 90 to 95% of the ferrocyanide can be oxidized in this manner. To complete the oxidation of the remainder, the soln. is electrolyzed between Ni electrodes, at a c. d. of 0.01-0.005 amp. per sq. cm., an initial e. m. f. of 7-8 v. and a final e. m. f. of 10-12 v. A porous diaphragm is used, and a small quantity of CaO is added to the anode compartment to neutralize the free $\text{H}_2\text{Fe}(\text{CN})_6$ formed there. H. STORITZ

Increasing the purity of common salt. THOS. B. BRIGHTON AND CARL M. DICE. *Ind. Eng. Chem.* 23, 336-9(1931)—Purification of solar salt made from Great Salt

Lake brine is described. Special attention was paid to elimination of odor and color

L. A. PRIDGON

Is solid-carbon dioxide manufacture as a kiln by-product practical? C. L. JONES
Pit and Quarry 22, 53-7 (1931).

E. H.

Chromite in 1929. LEWIS A. SMITH. *Bur. Mines, Mineral Resources of the U S*
 1929, Pt. 1, 203-29 (preprint No. 10, published March 5, 1931)

E. H.

Colloidal graphite as a protecting medium for automobile radiators. O STEINITZ.
Automobiltech Z 34, 236 (1931) — A suspension of a colloidal graphite of 0.000001 mm
 diam in water (about 1/10%) prevents the sedimentation of salts on the walls of the
 app and inhibits the corrosive action of the water on metal parts. A. A. BOENTLINGER

Revivification of bleaching earths. O ECKART *Seifensieder-Ztg* 58, 200-1
 (1931) — A revived earth recovers on the av 60% of its original bleaching power.
 The oil is first extd. by some solvent and the solvent and the earth are then roasted or
 burned.

P. ESCHER

The manufacture of bromine in France. M KALTENBACH. *Chimie & industrie*
 25, 543-55 (1931) — A detailed description of the process in use in French Br-mfg
 plants, dealing more particularly with the latest refinements in the processes for the
 manu and refining of the Br.

A. PAPIEAU-COUTURE

Properties of iodine. J LANZA. *Anales soc. españ. fis. quim.* 29, 221-34 (1931) —
 Aq I solns show marked supersatn phenomena, which explains many high soly re-
 sults. There is a max soly at 60° of 1 in H₂O or aq KI soln. In the system H₂O-
 KI-I contg an excess of the 2 latter components there is a reciprocal soly of KI for
 I, so that a soln satd with KI can still dissolve considerable quantities of KI when I
 is added, and reciprocally. In dil solns the molar relation KI:I is practically const.

E. M. SYMMES

An outline of the manufacture of carbon black. D. F. CRANOR AND L. J. VENUTO
Am. Paint J 15, No. 26, 2f-2 (1931) — The history, manu, uses and testing of carbon
 black are discussed briefly.

G. G. SWARD

Preparation of copper-chromium oxide catalysts for hydrogenation. RALPH CON-
 NOR, KARL FOLKERS AND HOMER ADKINS. *J. Am. Chem. Soc.* 53, 2012-3 (1931) —
 (NH₄)₂CO₃·H₂O (71 g) in 400 ml H₂O is added to a soln contg 50 g Cu(NO₃)₂·3H₂O,
 5.4 g Ba(NO₃)₂ and 77 g Cr₂(NO₃)₃·15H₂O in 575 ml H₂O, the ppt. washed twice with
 50-ml portions of H₂O, dried at 100-10° overnight, ground to a powder and heated to
 approx. 230°.

C. J. WEST

Utilizing C saturated with O [for manufacture of CO] (Ger. pat. 506,274) 21.
 Apparatus for carbonating solutions (U S pat. 1,799,354) 1.

Packing for autoclaves, etc. f G. FARBENING A-G (Otto Leuchs, inventor).
 Ger. 511,761, Jan. 9, 1925. The packing consists of fibrous material soaked in starchy
 fatty acids or derivs. The example mentions asbestos fiber soaked in glycerol and
 treated with the Na salt of a starchy fatty acid.

Nitric acid. E. I. DU PONT DU NEMOURS & Co. Ger. 522,167, May 7, 1929. See
 Brit. 336,638 (C. A. 25, 1953).

Use of hydrocyanic acid gas. SOC. FRANÇAISE DU GAZ SANOS. Fr. 697,923-4,
 Sept. 24, 1929. App. are described for emitting an alarm gas along with hydrocyanic
 acid gas to prevent poisoning of persons thereby.

Sulfuric acid. RICHARD ZEH. Fr. 697,884, June 25, 1930. See Ger. 509,703
 (C. A. 25, 778).

Removing fluorine compounds from gas for sulfuric acid manufacture. TAKESABURO
 IKEBE, TAITU OTANI, TATU KIKUTI AND MITSUI KOZAN KABUSHIKI KAISHA.
 Japan 89,600, Dec. 18, 1930. F compds are removed from SO₂ by passing through a
 layer of inorg. material, such as pieces of glass or bricks contg H₂SiO₄ or sand and then
 washing with H₂SO₄ (30° Be) contg alkali sulfites, sulfates or nitrates. Silica gel or
 K₂SiF₆ or the like is recovered from the washing liquor.

Ammonia synthesis. ALPHONS O. JAEGER (to The Selden Co.). Can. 310,208,
 Apr. 7, 1931. NH₃ is synthesized by passing a gaseous mixt. contg. H and N at an
 elevated temp. over a contact mass contg. at least one dild. "permutogenetic body,"
 at least part of the diluents being Fe-contg. products. Cf. C. A. 24, 1941.

Alkali nitrates. FRIEDRICH JOST. Fr. 698,404, July 4, 1930. An alkali nitrate
 is made from an alkali chloride, HNO₃ and NH₃ by using an alkali sulfate as intermediate
 product. The latter is transformed by HNO₃ into alkali nitrate and a soln. of alkali
 bisulfate in HNO₃. The alkali nitrate is sep'd and the soln. neutralized with NH₃
 and alkali chloride. The alkali sulfate formed is reutilized. Cf. C. A. 24, 1943.

Aluminum-alkali chlorides. METALLGES. A-G (Freiherr Conway von Gerswald, inventor) Ger. 522,031, July 5, 1923. Al is treated with a fused heavy metal chloride, e. g., $PbCl_2$, in the presence of an alkali chloride. The heavy metal is obtained as a by-product. An example is given.

Cyanogen compounds. GEORG BERING and EGON FLEISCHER Ger. 522,253, Dec. 8, 1922. A mixt. of a volatile N compd., a considerable excess of a volatile C compd., and a substantial amt. of H is passed at a dull red heat over a metal oxide, other than an oxide of a metal of the Pt group. Thus, a mixt. of NH_3 , 5, CO 15, and H 50 pts. by vol. may be passed over Cr_2O_3 at 600° , yielding HCN. Other suitable oxides are Al_2O_3 , ThO_2 , and ZrO_2 .

Nitrates from nitrites. FRIEDRICH A HENGLEIN (to I G Farbenind. A-G) U. S. 1,798,533, March 31. Nitrites such as those of K or Na are treated with O at temps. of $180-500^\circ$ in the presence of a catalyst of the group consisting of the alkali metal hydroxides, alk. earth metal oxides or their mixts. (at least one of the reaction materials being in the fused state).

Metaphosphates. I G FARBENIND. A-G (Otto Balz and Ernst Mische, inventors) Ger. 522,169, Nov. 29, 1929. Alkali or alk. earth metaphosphates are prepd. by heating the corresponding monoorthophosphates to about $400-500^\circ$ in the presence of inert solids which do not soften or melt at the reaction temp.

Separating potassium phosphates and aluminates. I G FARBENIND. A-G Ger. 522,168, Nov. 24, 1929. Solns. contg. phosphates and aluminates of K, with or without KOH or K_2CO_3 , are treated with NH_3 gas. The soln. seps. into two layers, the upper contg. mainly aluminate, and the lower contg. the phosphate with some aluminate. The layers are sepd., and the sepn. is completed by moderately dilg. the lower layer with water and repeating the treatment.

Phosphorus halides. N. V. ELECTRO-CHEMISCH. INDUSTRIE. (Henri C. J. H. Gelissen, inventor) Ger. 522,270, Sept. 26, 1929. See Fr. 664,830 (C. A. 24, 926).

Zirconium salts. DEUTSCHE GASGLÜHLICHT. AUER-GES. M. B. H. Fr. 698,193, June 28, 1930. Oxes. of Zr contg. ZrO_2 and SiO_2 are attacked or dissolved by alkalis in amt. such that for one mol. of ZrO_2 one mol. of alkali oxide is used. Na_2ZrO_4 and Na_2ZrSiO_4 are formed, and if SiO_2 is present in excess $Na_2ZrSi_2O_7$ is also formed.

Alumina purification. BASIL T. HORSFIELD (to Aluminum Co. of Am.) U. S. 1,798,261, March 31. For obtaining substantially pure alumina from alumina partially purified by reaction of a reducing agent at high temp. with assoed. oxide impurities such as Fe, Ti, Si and Ca oxides, the partially purified alumina in subdivided form is passed countercurrent through a stream of acid reagent of gradually increasing strength so that some unreduced oxide impurities are dissolved, and is then passed countercurrent through a stream of wash water contg. acid in quantities decreasing to zero.

Ammonium sulfate. WM. G. ADAM Fr. 697,657, June 19, 1930. Crystals of $(NH_4)_2SO_4$ which have little tendency to agglutinate on standing are made from an aq. soln. contg. a small proportion of a free acid and a small proportion of an appropriate metallic salt favoring the formation of elongated crystals, e. g., $Fe(SO_4)_2$. Cf. C. A. 24, 5048.

Ammonium sulfate. PATENTVERWERKUNGS A-G "ALPINA" (SOC. ANON. POUR L'EXPLOITATION DE BREVETS "ALPINA" PATENTE EXPLOITATION CY. "ALPINA" LTD.) Fr. 698,485, Jan. 31, 1930. $(NH_4)_2SO_4$ is made by neutralizing H_2SO_4 with NH_3 in 2 or more steps, the lye being cooled between each step. An app. is described.

Barium hydroxide. FRIEDRICH RÖSBERG and GUSTAV CLAUS (to Rhenania-Kunheim Verein Chemischer Fabriken A-G) U. S. 1,799,990, April 7. Ba silicate is subjected to fine grinding in the presence of water to cause hydrolyzing of the finely ground material. Cf. C. A. 24, 5115.

Calcium cyanamide. NIKODEM CARO and ALBERT R. FRANK. Fr. 698,231, June 30, 1930. $CaCN_2$ and other cyanamides are prepd. free from dust by treating the ground carbide in the form of pieces or briquets. Cf. C. A. 25, 1041.

Bleaching powder. SOC. ELETTRICA ED ELETTROCHIMICA DEL CAFFARO, ACHILLE CARUGHI and CARLO PAOLONI. Austrian 121,976, Nov. 15, 1930. Bleaching powder contg. more than 40% of active Cl is prepd. by chlorinating a suspension of $Ca(OH)_2$ in CCl_4 or other inert non-solvent for $Ca(OH)_2$ and $CaOCl_2$. The $Ca(OH)_2$ is ground in the inert liquid before, during or throughout the reaction, which is effected at a temp. above 40° . When the reaction is finished, the mixt. is cooled to below 20° and the product is dried. App. is described.

Bleaching powder. I G FARBENIND. A-G Fr. 698,361, July 3, 1930. See Ger. 508,167 (C. A. 25, 781).

Bleaching lye. I G FARBENIND. A-G Fr. 698,753, July 8, 1930. A lye for

bleaching and other purposes contains Na_2PO_4 and sol glass. It is made from phosphoric Fe contg S₁ (obtained in the electrothermal manuf. of P) by reaction with Na_2CO_3 and air, leaching the product with water and evapg to dryness.

Lime from limestone. GLENN H. NILES (one-half each to Doherty Research Co and Arthur E. Truesdell) U. S. 1,798,802, March 31. A charge of limestone is heated in a kiln under a subatm pressure which varies in different portions of the kiln. Heat for burning is supplied by means of a current of hot gases which are passed upwardly through the charge, then removed from the top of the charge and reintroduced at a controlled rate and pressure into the lower portion of the charge admixed with heating gases in sufficient volume to maintain lime-burning temps, and the amt. of air employed is limited to that which is supplied at a controlled rate and pressure to produce the heating gases. App is described.

Iron oxide and sulfur dioxide. HORACE FREEMAN (to The Sulphide Research Corp., Ltd) Can 311,130, May 5, 1931. Iron sulfide ore is passed in finely divided form through a zone in which it is preheated to ignition temp. by contact with hot furnace gases, and through a second zone in which it is heated to fusion temp and burned with a deficiency of gaseous oxidizing agent whereby most of the S is oxidized and through a third zone in which it is further heated to a state of fusion and is burned with an excess of gaseous oxidizing agent whereby substantially all the Fe and the residual S are oxidized, and finally through a fourth zone in which the fused particles are chilled sufficiently to avoid agglomeration.

Potassium phosphate. ROBERT D. PIKE U. S. 1,799,479, April 7. In making K_3PO_4 from KCl and H_3PO_4 in a single furnace operation, a mixt. of KCl is furnaced with sufficient H_3PO_4 to form the tri-K phosphate and with sufficient H_2SO_4 to drive off substantially all the HCl, and a carbonaceous reducing agent such as coal or coke is added to complete the conversion of the furnace contents. An arrangement of app. is described.

Sodium sulfate from natural deposits. ARTHUR T. W. WARREN (one half to Emanuel Johnson). U. S. 1,798,993, March 31. The natural deposit is covered with fresh water which is then subjected to natural temp conditions of day and night so that the water is heated during the day to collect the Na_2SO_4 and cooled during the night to ppt. the material, and the warmed water is led during the day to a collection vat.

Purification of carbon dioxide. RALPH H. MCKEE Can 310,569, April 21, 1931. CO_2 contg low-boiling hydrocarbons is purified by cooling under a pressure of approx. 400-1000 lbs per sq in to provide a density of the CO_2 of approx. 5-60 lbs per cu. ft., and contacting with activated C, silica gel, solid paraffin in granular form or a heavy odorless oil such as refined white petroleum oil.

Apparatus for making solid carbon dioxide of high density. MASATOSHI OHKOTI, YOSIYOSI OYAMA and INSTITUTE OF PHYSICAL AND CHEMICAL RESEARCH, Japan 89,547, Dec. 13, 1930.

Gaseous sulfur compounds. KOLOMAN VON SZOMBATHY, KORNEL KELL and PAUL SCHMITZ. Fr 698,165, June 23, 1930. Ores contg. a metal and S are mixed with C and, while heated to incandescence, are exposed to a current of gas contg. steam. SO_2 as well as H_2S is obtained if an excess of air is mixed with the steam. Fe_2O_3 may be added as a catalyst. Fr 698,166 describes an app for absorbing H_2S and SO_2 in gases by means of a finely divided concd soln of polythionates sol in water, e. g., pentathionates or hexathionates of alkalis or alk earths. The S which seps is extd from the soln and dried, while the soln freed from S is returned to the reaction vessel. The process is accelerated by adding to the soln small amts of alcs, e. g., 1-2% of glycerol, and keeping the temp at 45-65°.

Continuous calcination of gypsum. ALVA W. TYLER U. S. 1,798,857, March 31. Finely divided raw gypsum is moved through a chamber in a stream of superheated steam at a suitable pressure; the superheat of the steam over its temp. of condensation at the pressure is sufficient to impart to the gypsum the heat necessary for its calcination without reduction of the steam to its temp of condensation. App. is described. Cf. C. A. 24, 1720.

Catalytic reduction of carbonyl compounds. ALPHONS O. JAEGER (to The Selden Co.) Can 310,207, Apr 7, 1931. Carbonyl compds are catalytically reduced by reaction with reducing gases in the presence of a contact mass contg a dild "permuto-genetic body."

Hydrogen. JOHN S. BEEKLEY (to DuPont Ammonia Corp). U. S. 1,799,452, April 7. A mixt. of steam and CH_4 in the vol. ratio of at least 2:1, at a pressure of 1-50 atm. and at a temp. below 600°, is subjected to the combined action of a Ni cata-

lyst and magnesia (the latter being present in the proportion of at least one mol of MgO per mol of CH_4 in the gaseous mixt.) *Cl. C. A.* 25, 1345

Hydrogen. HEINRICH BOMKE *Ger* 516,843, Oct. 17, 1926 *Il.* is prepd. by passing water gas or similar mixts over a MgO catalyzer. CO_2 absorbents may also be mixed with the catalyzer. A suitable catalyzer mixt. is given by burned lime contg MgO .

Hydrogen. PATENTVERWERTUNGS A-G "ALPINA" (SOC. ANON. POUR L'EXPLOITATION DE BREVETS "ALPINA" PATENTE EXPLOITATION CY. "ALPINA," LTD). *Fr.* 698,484, Jan 31, 1920 *Il.* is obtained by washing coke-oven gas with water under pressure. The gas which is first liberated from the water on reducing the pressure contains 30-35% of *Il.* and is returned to fresh unwashed gas.

Hydrogen and other gases. GEORGE F. JAUBERT *Fr* 698,706, Oct. 9, 1929. Gases such as *Il.* are prepd with or without pressure by interposing between the liquid (caustic alkali or other) and the solid substance (*Ss.* or other) to be attacked by this solid, a partition or pocket made of appropriate material attackable itself by the reaction liquid but with less rapidity, to permit the charging of the reservoir which is to contain the gas under pressure, and the closing of it before the reaction takes place.

Graphite and hydrogen. I. G. FARRENING A-G (Hans Dahr, inventor) *Ger.* 516,901, Oct. 4, 1929. Graphite and *Il.* are produced by the decompos. of hydrocarbons in metal baths. App. for leading the hydrocarbons into the molten metal at suitable temps and at a suitable velocity is described. In examples, hydrocarbons are led into fused Fe and fused Cu at 1300° .

Active carbon. AKT-GES. FÜR STICKSTOFFDARST. (Theodor Geis and Walter Pechtold, inventors) *Ger* 516,881, June 6, 1923. Highly active C is prepd by heating flue ashes of all varieties (e. g., from coal, lignite, peat and wood fuel) to about 1000° , and not exceeding 1050° . *Cl. C. A.* 24, 1710

Apparatus for production of activated carbon. WM. M. WILLIAMS (to Activities Ltd.) *U. S.* 1,799,827, March 31. Carbonaceous material such as coal dust is injected in a tangential direction into a combustion chamber in the form of an elongated furnace of rounded cross section. *Cl. C. A.* 24, 1710

Apparatus for producing carbon black by partial combustion of hydrocarbons. CHAUNCEY MATLOCK (to Monroe Louisiana Carbon Co.) *U. S.* 1,798,614, March 31. Various structural details are described.

Adsorptive charcoal. BEARNARDINO OCLIFETTI *Fr* 698,752, July 8, 1930. Charcoal having the power of adsorbing gases and vapors, decolorizing charcoal and charcoal resembling lamp black are prepd by calcining, in the absence of air, residues obtained by evapng solns. of org. substances obtained by treating vegetable materials with hot solns. of $NaOH$ or KOH , washing the residue with water, treating it with strong mineral acids (H_2SO_4 , HCl , etc.), washing with water, drying and grinding.

Phenol-formaldehyde condensation products. EDWARD S. HOLE *U. S.* 1,799,816, April 7. A non-sticking condensation product is made by condensing carbohic acid and CH_2O by boiling in the presence of an alkali salt and dilg. the material with water while it is still boiling, cooling the dil. mixt., and subsequently adding more water contg. a Cr salt. The product is suitable for coating or impregnating.

White phenol-formaldehyde condensation product. TYOJI ARIYAMA and TYUTARO SIBUYA *Japan* 89,615, Dec. 19, 1930. A mixt. of ammonium or alkali borate, sol. salts of Sn, Ti or Sb, $HCHO$ and $PbOH$ is heated to give a white condensation product.

Condensation products from formaldehyde. LUDWIG ORTHNER *Fr.* 697,678, June 19, 1930. A mixt. of glycolic aldehyde, trioses and tetroses is prepd. by heating an aq. soln. of CH_2O with the addn. of an alk. reacting condensing agent such as $Pb(OH)_2$ until the reducing power of the reaction liquid for cold Fehling soln. has reached a max.

Gelatinizing urea-formaldehyde condensation products by heating with sucrose. OSCAR A. CHERRY (to Economy Fuse and Mfg. Co.) *U. S.* 1,799,954, April 7.

Artificial substances. I. G. FARRENING A-G *Fr* 697,693, June 19, 1930. A mixt. of vinyl compds. or a mixt. of one or more vinyl compds. is polymerized with other compds. capable of being polymerized and contg. the group $=C=CH_2$, if necessary in the presence of a catalyst and a solvent and adding fibers and emollients. Thus Et acrylate and vinyl acetate are polymerized together. Other examples are given. *Cl. C. A.* 24, 4128

Molded articles. JONAS H. SCHMIDT (to The Bakelite Corp.). *Can.* 309,680, Mar. 24, 1931. Molded articles are made by shaping a fusible resinous substance of glycerol-polycarboxylic acid type in a mold, converting it while under pressure into an infusible resin, removing it from the mold and hardening it by further application of heat. *E. g.*

a mixt. of 200 parts by wt. of glycerol with 400 parts of phthalic anhydride, with or without fillers or coloring materials, is allowed to react at 140–200° to incipient gelation, and then poured into molds and heated under pressure until gelation occurs and the resin passes into the infusible state (the essential requirement is that the chem. or phys. change from fusibility to infusibility should take place in the mold and preferably under the molding pressure). The shaped article does not tend to lose its form or alter its dimensions on further heating.

Plastic materials. C. F. BOHRINGER & SOHN G. M. B. H. Fr. 698,321, July 3, 1930. The elasticity of plastic materials and objects made therefrom having a basis of cellulose derive is increased by producing, either by addn., subtraction or substitution, a layer free from or poor in plasticizing agents. Thus, the articles may be dipped in ether or sprinkled with a soln. of ethylcellulose in CaH_2 .

Plastic materials. SOCIÉTÉ ANON. DE DÉCORATION ARCHITECTURALE ET MOBILIÈRE. (Gustave Moncany, inventor) Fr. 697,435, June 14, 1930. A plastic material capable of being molded is made by mixing plaster, fibers and asbestos powder and incorporating albumin or dextrin, molding, baking, dipping into a bath of water and glycerol, baking again, dipping into synthetic resin in alc. and baking. Fr. 697,436 describes a plastic material from asbestos fiber and paper beaten up with a synthetic-resin varnish to which metallic fibers may be added.

Plastic materials. THOMAS & CIE G. M. B. H. Fr. 698,213, June 30, 1930. Pentamethylenediaminesulfone is used as a constituent of plastic materials, either as the principal constituent or as a flux or softening agent.

Porous masses from gypsum. RHEINOLD & CO., VEREINIGTE KIESELGLAS- UND KORKSTEIN-GEF. Ger. 522,139, Mar. 6, 1929. Gypsum is mixed with relatively large amts. of substances which increase the porosity of the product, e. g., sugar or glue, and to compensate for the delaying action of these addns. on the binding of the product, there are added also accelerators of the binding, e. g., a sulfate or a nitrate. The products are useful as heat insulators, filter stones, etc.

Unbreakable transparent sheets. LOUIS C. F. SCHMIDT Fr. 698,411, July 5, 1930. Unbreakable transparent sheets are made from condensation products of phenols with aldehydes, having transparent sheets of cellulose as interior or exterior supports.

Silica gel. KOZO IKAWA, SUTKA ISIKAWA and MITSUI KOZAN KANUSHIKI KAISHA, Japan 89,500, Dec. 18, 1930. Silica gel is manufd. by introducing Cl_2 or a mixt. of Cl_2 and a gas which does not react with alkalis into water-glass soln. The product is washed with H_2O and dried. The filtrate contains about 80% NaOCl .

Adsorbents from gels. FRITZ STOEHLER (to I. G. Farbenind. A.-G.), U. S. 1,798,766, March 31. Part of the water-sol. impurities are removed from a gel such as that of silica by washing, the material is partially dried, washed again and then completely dried.

Material for drying gases. KOZO IKAWA and MITSUI KOZAN KANUSHIKI KAISHA, Japan 89,593, Dec. 18, 1930. Anhydrous MgSO_4 is mixed with porous or colloidal material like CaSO_4 , cement, silica gel, activated C, Japanese acid clay, powdered glass, Al_2O_3 , etc. The product absorbs much H_2O , and can be used repeatedly by drying with heating.

Dehydrating material suitable for use in show casings, buildings, etc. HOWARD H. BAKER, U. S. 1,798,802, March 31. A solid molded mass is formed of CaCl_2 99% and 1% of metallic particles such as iron filings. Sawdust may be added.

Adhesive. I. G. FARBENIND. A.-G. Fr. 697,874, June 25, 1930. A cold adhesive is made from condensation products of urea, thiourea or their derivs. with aldehydes or their polymers, water, and, if necessary, acids, salts or substances liberating acids and substances which increase the consistency of the adhesive.

Adhesive for gummed paper. FREDERICK W. HUMPHREY (to Mid-States Gummed Paper Co.) U. S. 1,799,797, April 7. Paper or fabric is coated with an adhesive comprising a mixt. of 2 resinous gums (such as pontianak and copal) having different rates of soly., and a small percentage of non-drying vegetable oil such as castor oil.

Cement. CONRAD MAYER, Fr. 698,010, May 13, 1930. A cement for repairing porcelain, etc., consists of a mixt. of flowers of chalk, freshly slaked lime, milk casein, quartz powder, glass powder, asbestos powder and coloring material.

Cement. MYRON E. DELANEY (to The Bakelite Corp.), Can. 310,617, April 21, 1931. A cement comprises about equal parts by wt. of a wax, such as chlorinated naphthalene, ester gum and about 5% of rubber latex. This cement is used for securing glass to glass or metals, wood to wood, etc.

Clutch-facing fabrics treated with lead oleate. SAMUEL P. HOWE, U. S. 1,798,605,

March 31 A fabric such as that for use as a clutch facing is treated with Pb oleate Mica, Pb wire, Pb linoleate and asbestos also may be used in the material

Composite sheet material auth as fiber and bakelite products. LOUIS T. FREDERICK (to Fibroc Insulation Co) U S 1,799,505, April 7. Fibrous particles such as comminuted scrap canvas are felted by the action of suction operating through a pervious backing strip such as textile fabric which forms a part of the finished sheet, and the latter is treated with a suitable resin such as reactive "bakelite" compn App is described.

Impregnated bands. BAKELITE CORP. Fr 698,358, July 3, 1930. Fibrous materials and plastified resinoids are mixed in the presence of water The mixt. is formed into bands which are dried, superposed and united by compression and heat

Emulsions. FREDERICK W ATACK. Can 309,760, Mar. 24, 1931 Stable, fluent emulsions are made by emulsifying a chlorinated hydrocarbon with water and Turkey red oil in the presence of a colloidal substrate, e g, 67% C_2H_5Cl , 4.5% Turkey-red oil, 2% bentonite, 26% water, 0.45% ammonia (d 0.880).

Sealing composition. ERNEST R. HANSON and MYRON E DELANEY (to Bakelite Corp) Can 309,681, Mar 24, 1931 A sealing compn is prepd. by melting a wax-like halogenated naphthalene and dissolving gilsonite therein This compn. is adapted for use in the manuf. of fixed elec condensers

Potting compositions. MYRON E DELANEY (to The Bakelite Corp). Can 311,044, May 5, 1931 A potting compn comprises 25-40% of a wax-like chlorinated naphthalene, 20-30% of gilsonite and 30-50% of montan wax and has a flow point ranging from 80° to 100° and a viscosity of less than 100 at 130°. This compn. is used on fixed elec. condensers, dry batteries, etc.

Arc-resistant composition. FRAZIER GROFF (to The Bakelite Corp) Can 309,679, Mar 24, 1931 An article resistant to the carbonizing action of an elec. arc is prepd. by causing phenol to react with tung oil and a methylene-contg agent in proportions to yield a resinoid product, admixing a cellulose filler therewith, partially baking the mixt. and molding under heat and pressure. Cf C A 25, 1046

Wetting, etc., agents. H TH BÖHME A-G Fr 698,380, July 4, 1930 Wetting cleaning, emulsifying, etc., agents are prepd by sulfonating amines or hydroxyamines of the aliphatic series contg more than 8 C atoms Examples are given of the sulfonation of octadecenylamine. Cf C A 25, 2532

Wetting, etc., agents. CIGMISCHE FABRIK FORM SANDOZ. Fr. 697,350, June 13, 1930 Monoaryl esters of glycerol and glycol are used along with carboxylic or sulfonic acids as wetting and emulsifying agents.

Wetting, foaming, dispersing and emulsifying agents. I G FARBENFABRIK A-G (Karl Daumler inventor) Ger 521,005, Jan. 26, 1929 Quaternary ammonium salts suitable for the above purposes are prepd. by converting $Et_4NC_2H_5OH$ into a quaternary salt by reaction with a alkyl halide, and then esterifying the OH group by means of a fatty acid chloride contg at least 5 C atoms in the mol Thus, a mixt. of $Et_4NC_2H_5OH$ and $PhCH_2Cl$ may be kept at 80-100° until the $PhCH_2Cl$ has disappeared, and oleic acid chloride then run in.

Wetting, cleansing and emulsifying agents. I G FARBENFABRIK A-G. Fr. 698,637, July 8, 1930 There are introduced into org compds aliphatic, cycloaliphatic, mixed aliphatic-aromatic or aromatic radicals contg alc. OH groups, ethylenic double bonds or sulfonic ester groups and also other reactive groups (The esterification of COOH groups of fatty acids contained in fats and oils and their sulfonic esters is excepted.) The products are converted into sulfonic esters or their salts if they do not already contain sulfonic ester groups Thus a mixt. of BuOH contg HCl and the product obtained by the partial chlorination of fish oil are boiled under reflux and the product is sulfonated. The Na salt of the product has a high wetting and emulsifying power Other examples are given. Cf C A 25, 2824

Detergent. E KOLOZSY Hun. 102,126, June 11, 1930 An aq dispersion of natural or artificial latex is mixed with ethereal or other org perfume and inorg or org filling materials.

Waterproofing agents. CHARLES H THOMPSON and WM J MCGIVERN Fr. 697,866, June 24, 1930 A product useful for waxing, waterproofing, etc., consists of an emulsion prepd. by adding to molten wax an alk. compd., e g, NaOH, KOH, Et_4N or ammo compds, and with or without a protein soln. or a colloidal substance. The emulsion may be mixed or combined with soap and used as a supplying agent.

Anti-freeze solution. CHARLES V. McAVOY (to McAvoy Products Co) U. S. 1,798,547, March 31 See Can. 299,764 (C. A 24, 3093).

Freezing mixtures. ERN. HUSMANN. Ger 518,880, Aug 10, 1929. A freezing

mixture which maintains a low temp for a considerable time is obtained by adding ice to an aq. salt mixt below a layer of specifically heavier non miscible liquid such as CCl_4 . The salt soln may contain glycol and the salt is preferably in granular condition.

"Artificial" (imitation) snow. MAX C RAUMANN U S 1,800,187, April 7. A white readily crystallizable org acid such as benzoic and salicylic acid is mixed with a light-colored adhesive material such as sugar, potato flour and alum.

Stiffening sheet material such as toe boxes of shoes. WILLIS A BOUGHTON (to Bennett Box Co) U S 1,799,919, April 7. A material such as cloth, leather or paper is treated with a sol silicate and glycerol Cf C A 24, 2818.

Porous water-laid felt. HENRY P SHOPNICK (to Arden Box Toe Co) U S 1,799,931, April 7. A porous water-laid felt is made with a preponderating proportion of long fibers bonded together with a smaller proportion of the gelatinized portion of untanned animal skins. This product is suitable for use as a box toe stock.

Planographic printing process. VALENTIN DIETRZ (to American Multigraph Co) U S reissue 18,013, March 31. Reissue of original pat No 1,741,758 (C A 24, 1191).

Stencil sheet. SHUNJIRO HOKUJI U S 1,799,793, April 7. A base of fibrous material such as Japanese yoshino paper is coated with a compn comprising a gelatinous org colloid such as gelatin dispersed in water in admixt with naphthalenic acids Cf C A 25, 1938.

Transfer. WINTHROP S. LAWRENCE (to Kaumagraph Co). U S 1,799,754, April 7. A paper base permeable to steam and free from any filler which would prevent the passage of steam is provided with a surface marking comprising a dye sol in water and embodied in a surface film such as a gum arabic compn which is sol in water, so that if steam is forced through the base it can act directly upon the film and dye Cf C A 25, 605.

Cigaret tips. I G FARDENING A-G (Walter Becker, inventor). Ger 516,710, Oct 6, 1929. Use is made of an under layer of resinous polymerized vinyl derivs and a water proof over layer. In an example, the over layer is the product of cellulose acetate, tricresyl phosphate and acetone, and the under layer of bronze powder and vinyl resin.

19—GLASS, CLAY PRODUCTS, REFRACTORIES AND ENAMELED METALS

G B DARTON, C H. KERR

Opacification of glass by blast lamp. M A BESHORODOV AND M F. SHUR. *Keram. i Steklo* 6, 466-70 (1930). —The authors describe tests made to ascertain the dependence of the velocity of glass opacification on (1) the chem compn of the glass; (2) the kind of flame used, (3) the method of working the glass, and (4) the quantity of cullet used in the glass. Fifteen kinds of glass were treated with reducing and oxidizing flames. The results are tabulated and show that the velocity of opacification depends on the chem. compn. of the glass. Glass contg CaO and alkalis becomes opaque most rapidly; glass contg ZnO becomes opaque less rapidly. Glasses having a high content of Al_2O_3 and B_2O_3 become opaque only with reducing flames. With oxidizing and reducing blast flames they do not become opaque at all. The presence of MgO in glass seems to aid opacification. Glass contg Al_2O_3 (9.3%) does not become opaque when worked with any kind of flame. Glass contg PbO becomes opaque less rapidly in a reducing flame than in an oxidizing flame. A high quantity of SiO_2 aids opacification. All glasses become opaque more rapidly in a reducing flame than in an oxidizing flame. The opacification of glass is a reaction taking place on the surface of glass through the building of new formations by the red hot gas particles. Some formations have the character of crystals, others that of amorphous efflorescences. The quantity of cullet does not influence the opacity of glass while being worked in a blast lamp flame. A series of other salt compds besides NaCl was tested to det their influence on the opacity of glass. KCl removes the opacity completely and more rapidly than NaCl . Br and I salts of K and Na also remove opacity, although the glass is not as transparent as when KCl is used. Bone acid removes opacity in glasses contg B_2O_3 , Pb_2O_3 and ZnO also remove opacity. M V. KONDOIDY

Glass technology and industrial progress. W. E. S. TURNER. *J. Soc Glass Tech.* 14, 173-84P (1930). —A general discussion by T. and others. H F. KRIEGE

Progress in glass technology under the auspices of the glass delegacy. ANON. *J. Soc Glass Tech* 14, 307-13T (1930). —The annual rept. of the Glass Research Delegacy includes a list of the departmental publications. H F. KRIEGE

International glass standards. ANON. *J. Soc. Glass Tech.* 14, 158-9 P (1930).—The Deut. Glastechnische Ges. and the Soc. of Glass Technology agreed upon the following standards: (1) For general comparative purposes the thermal expansion of glass shall be detd. for a mean temp. of 50° over a range of 25° = 5° to 75° = 5°. The specimens shall be 5.0 mm. thick and normally free from strain as detd. optically. The symbol of mean linear coeff. of thermal expansion shall be α_{100} . With measurements made at high temps. a test piece shall be used only once and the rate of heating shall be 2° per min. (2) The viscosities and the corresponding temps. important in the characterization of glass are: (a) 10¹⁰ c. g. s. units—furnace working temps., (b) 10⁸ c. g. s. units—temp. of the lower limit of the working range, (c) that at which a break occurs in the temp. property curve. This is given as 10¹⁰ c. g. s. units where the property considered is elec. cond. (3) The symbol T_g shall replace all other conventions to denote the transformation temp. of glass. The max. point on the complete thermal-expansion curve for glass, namely, its softening temp., shall be denoted by M_s . H. F. KRIEGER

Thermal endurance of glass. W. J. A. WARREN. *J. Soc. Glass Tech.* 14, 313-23 T (1930).—The thermal-endurance detn. proposed by Gould and Hampton is critically examd. Their formula is $\theta \sqrt{t}$, where θ is the min. temp. difference causing a tall conical beaker (200 cc. capacity) contg. hot wax to fracture when plunged into cold water and t is the max. thickness in the bottom of the beaker along the course of the fracture. W. finds it necessary to specify that the max. thickness shall be within the concave part of the bottom in order to have concordant results. Circumferential fractures give very slightly higher and more consistent values than transverse cracks. Uneven distribution of the glass has no apparent effect where the ratio of max./min. thickness is 1.5-2.0. The thermal endurance of beakers of sheet glass compn. was found to be deteriorating several months after manufacture. The best value of α in the formula, thermal endurance, $B = \theta t^\alpha$, was found to be 0.36-0.40, where the beakers average 14 mm. in thickness. A more satisfactory expression is $B = \theta + (t-1)A$, where A is a const. for any definite compn. which is deduced from a few beakers tested by putting $A = B/3.75$. H. F. K.

A provisional standard test for the thermal endurance of glass. ANON. *J. Soc. Glass Tech.* 14, 160-3 P (1930).—The test procedure accepted by the Glass Standards Comm. of the Soc. of Glass Tech. uses a tall conical beaker (200 cc.) 140 mm. high, 72 mm. diam. base and 49 mm. diam. top as the standard article. The thickness of the beakers should be as close as possible to 1 mm. The beakers are tested from 2 to 48 hrs. after manuf. to avoid irregularities due to weathering, etc. Beakers of satisfactory standard of annealing are selected. Paraffin wax is melted in each beaker to form a layer 4-5 cm. deep and heated to a few degrees above the temp. desired. When the immersed thermometer reads the desired temp. such as 75°, the beaker and contents are plunged into water at 15° = 1°. If fracture does not occur the process is repeated at the next higher heating temp. The testing should begin at a temp. low enough to insure a no. of chillings before rupture occurs. The heating temps. increase in steps of 5°. The difference between the temp. of the chilling bath and the temp. to which the beaker was last heated before fracture occurred is recorded, likewise the thickness of the beaker along the fracture and the max. thickness in mm. Any beaker having a min. thickness less than 1/2 its max. thickness along the fracture is rejected. The results of 10 beakers whose max. thickness is 0.9-1.0 mm. are averaged to obtain the thermal endurance of the glass according to the equation: thermal endurance = temp. difference $\times \sqrt{\text{thickness in mm.}}$. H. F. KRIEGER

Calculation of the tensile strength of glass. O. KNAPP. *Glashütte* 60, 364 (1930). *J. Soc. Glass Tech.* 14, 290-1A.—For the purpose of calcn. of the tensile strength, it is assumed that this is an additive property dependent on the silicates present in the glass, and that when the SiO₂ content is 50-66.6 mols. % the oxides are present as monosilicates and the excess SiO₂ is present as free oxide. Similarly, when the SiO₂ content is 66.6-75 or 75-80 mols. % the oxides are present as di- or tri silicates. The tensile strength of such a glass is then readily calcd. by use of factors, deduced from the data for tensile strength published by G. Gehlhoff and M. Thomas (*C. A.* 22, 1448). A table shows the value for the tensile strength of 22 of Gehlhoff and Thomas' glasses as calcd. (a) by A. Winkelmann and O. Schott's factors, (b) by Gehlhoff and Thomas' method and (c) by the above-mentioned factors. The percentage deviation of each of these from the observed value is shown. Methods (b) and (c) are fairly accurate but (a) gives rather high values. G. G.

Critical investigation of the methods for the determination of the resistance to

chemical attack of glasses ERNST REXER. *Keram. Rundschau* 38, 387-90, 421-5, 464-5, 506-7, 546-8, 560-4, 590-3, 624-7(1930)—The methods now in use for testing the attack of water on glass were critically investigated. Methods using crushed samples, the Mylius "surface samples" and autoclave samples were tried. Relations between soly and size of sample, fineness of sample, character of surface and amt of strain were obtained. H F INSLEY

Effect on the properties of a soda-lime-silica glass of repeated melting in platinum. The x-ray pattern. W E S TURNER. *J Soc Glass Tech* 14, 351-3T(1930), cf C A 25, 788—No change in phys. properties was observed unless a chem. change was produced by the repeated meltings in Pt. H F K

Development and trend of the mechanical production of white hollow glass ware. A WENDLER. *J Soc Glass Tech* 14, 353-68T(1930). H F K

The supposed diphasic nature of glass. F W PRESTON. *J Soc Glass Tech* 14, 349-50T(1930)—Double refraction has been noted in glass allowed to cool under pressure. P suggests that this phenomenon may be due to the presence of ultramicroscopic crystals with some orientation. H F K

Use of borax in glass manufacture. EDWARD MEIGH. *J Soc Glass Tech* 14, 164-72 P(1930)—A general discussion by M. and others. H F KRIEGER

The development of tank-block production. HANS HIRSCH. *Keram. Rundschau* 38, 291-5, 312-5, 343-6(1930), cf C A 24, 5121—A no. of commercially made tank blocks were divided according to Al_2O_3 content into 2 classes. Besides chem. analyses, phys. tests for sp. gr., porosity, compressive strength, abrasive hardness (resistance to wear), resistance to thermal shock, resistance to molten glass attack and to Na_2CO_3 attack and refractoriness were made. Although results show ambiguity due to differences in methods of manuf. as well as compn., it is concluded that low porosity with only fine pores present, high d., uniformity in compn. and structure and high strength give the best tank block and that, other things being equal, blocks high in Al_2O_3 give better resistance to glass attack than those low in Al_2O_3 . H INSLEY

The thermooptical properties of glasses rich in silica. F RINNE. *Keram. Rundschau* 37, 772(1929), *J Soc Glass Tech* 13, 371A—The ns of glassy substances were detd. for wave lengths of 735-440 μ for temps. between -160° and 1000°. The following substances were used: (a) fused quartz, (b) hyalite (SiO_2 , 3.4% H_2O), (c) hyalite (SiO_2 , 4.84% H_2O), (d) opal (SiO_2 , 8.5% H_2O), (e) opal (SiO_2 , 9.73% H_2O). As the water content increased, the n was lowered. The dispersion was const. A max. was detected in the n temp. curve at 0° with glass (e). No sepn of water could be detected in an opal on cooling in liquid air. G G

Elimination of waste in the glass industry. GEO. H. MILES. *J. Soc Glass Tech* 14, 185-95 P(1930)—General. H F KRIEGER

Continuous kilns for the burning of clay products. J WILLIAMSON. *Trans. Ceram. Soc. (England)* 30, 21-32(1931), cf C A 25, 789. H F KRIEGER

Manufacture of bricks for road paving in the U. S. A. and Holland. G A HODSON. *Trans. Ceram. Soc. (England)* 30, 37-45(1931). H F KRIEGER

Oil fuel and the brick industry. I LUBBOCK. *Trans. Ceram. Soc. (England)* 30, 68-80(1931), cf C A 25, 1066—To offset the higher cost of oil over coal 3 economies are claimed for oil fuel: (a) reduction of time of burning, (b) elimination of underburned brick, (c) saving labor costs for coal and ash handling. Various types of fuel-oil burners and oil-fired kilns are described. H F KRIEGER

Use of garbage slags for manufacture of slag bricks. R GRUEN. *Städtereinigung* 22, 71-8(1930)—A description of the new practice of refuse disposal at Cologne. Chem. analyses of garbage slags and their phys. properties and the compn. and test data on the phys. properties of brick manufd. from this slag are given. E. I. S.

Advances in the production of hollow tile. ALFRED HIELSCHER. *Tonind.-Ztg* 55, 379-82, 420-3(1931)—Modern machinery is described. B A SOULE

Method for determining the workability of clay and ceramic masses with the ball plastimeter. WILLI M. COHN. *Tonind.-Ztg* 55, 238-40(1931)—An app. is described for measuring the workability of clay mixes. It consists of a ball 20 mm in diam. attached to a plunger which is surmounted by a flat plate. The ball is made to rest on the mud and weights are placed on the plate. The depth of penetration of the clay into the mud is measured on an attached scale. The depth of penetration depends on the water content of the mud, the pressure on the plunger and the time the pressure is applied. Expts. were made with one clay, by varying one of these factors and holding the others const. The method is not claimed to be a measure of plasticity but rather a measure of the pressure component (yield value?) of the workability. R F F

Solid fuels in the fine ceramic industry. A. JÄESCHKE. *Keram. Rundschau* 37,

541-5, 561-2 593 G, 631-4(1929)—Temps required for burning different kinds of ceramic ware are given and various fuels are evaluated. R I S.

The effects of various gases in the kiln on the burning of bright gold for gilded ceramic wares. AKIRA NAKATSUCHI *J Soc Chem Ind, Japan* 34, Suppl binding, 67-9(1931)— SO_2 , H_2S , CO , CO_2 and H_2 are found to be injurious, but not at low temps. The best practice is to burn at 300° , sweep out the gases and heat to 750° . V F H.

Gas permeability of pyrometer tubes in relation to temperature. W. MRENA *Tonind Ztg* 55, 149(1931)—Contrary to general belief, the permeability of ceramic pyrometer tubes to gases was found to decrease materially with increased temp. This was explained by assuming that gases have a higher viscosity at high temp. Detns of the viscosity of CO_2 , N_2 and producer gas at 1220° confirmed this assumption. R F FERGUSON.

Acceleration of the drying of ceramic products. P P BUDNIKOV, G V. KUKOLEV AND L. L. MANDLERON *Tonind Ztg* 55, 401-11(1931)—The drying time may be lowered 50% by addn of electrolytes, especially 0.5% FeSO_4 or 0.25% $\text{Al}_2(\text{SO}_4)_3$. B A SOUL.

Silicates high in alumina. VLADIMIR ŠKOLA *Chem Listy* 25, 30-4, 53-61(1931)—The phase diagram for the system Al_2O_3 - SiO_2 is evolved and discussed in connection with the production of corundum, sillimanite, andalusite, etc. FRANK MARSH.

Some of the newer uses for silicon carbide. CHARLES McMULLEN *Trans Electrochem Soc* 59, preprint 6 pp (1931)—Important recent adaptations of SiC are recorded and discussed. In recuperators cylindrical tubes are used. In the "Carbo-radiant" combustion chamber a mixt of air and fuel is injected into a SiC box. The Thornhill Anderson furnace has an arched overhead combustion chamber made of SiC. Other applications include boiler furnace walls, heating elements, therapeutic lamps, resistors for radio equipment and lightning arresters. The physical properties of SiC are tabulated. C G F.

Grog. C R F THORPALL *Trans Ceram Soc (England)* 30, 1-18(1931); cf. C A 25, 183.—The best fireclay grog is made from raw lumps calcined at high temps and crushed to have a splintery shape. To make low porosity brick it is best to use grog with high porosity. H F KREGE.

Refractory materials. Observations on the behavior of refractory materials in continuous vertical retorts. A T GREEN *Inst Gas Eng* 1930, Communication No 22, Pt 3, 9-20, *Gas Eng* 47, 689 91(1930), *Gas J* 192, 720-4.—Most cases of "holing" were found to originate in the interior of the retort rather than in the combustion chamber. Foreign substances such as Fe compds and alkali from coal, cause superficial disintegration or corrosion. The reducing atm prevailing in the retort accentuates the slagging action of iron slags. The viscosity or stickiness of the slag at the surface is important in its influence on the passage of the charge through the retort. High-alkali slags cause augmented stickiness for alkali silicates are low melting. Material in the zone 6 to 12 ft from the top of the retort often peels off in thin layers. Examin of these flakes showed a very slight increase in Fe content. The alkali content increased from 0.52% to 0.73%, the sp gr decreased from 2.36 to 2.32. Flaking often takes place during the scurfing operation. The surface in contact with the scurf has probably been weakened at definite planes by (a) depositions and impregnations, (b) constitutional changes in the SiO_2 in the surface layers resulting from previous high temp scurfings, (c) a general spalling tendency of a SiO_2 product. Joints and jointing cements are considered. A H DIERKEA.

Refractory materials for electric furnaces. II. Copper and its alloys, including bronze, gun metal, brass, copper-nickel and phosphor-bronze. ALFARO B SEARLE *Metal Ind (London)* 38, 333 5(1931) cf C A 25, 1650.—The most important types of elec furnace for melting copper, brass, bronze, etc., are the direct arc, the indirect arc, the induction and the granitic resistance. The direct arc type furnaces are suitable only for bronzes and alloys contg less than 6% Zn, while the indirect arc types are applicable to all Cu alloys. The Pb content of the alloy detrs the type of lining—magnesite being preferable when Pb is present, though Zn may react with a MgO lining. In the Ajax-Wyatt furnace for melting yellow brass, a lining consisting of a refractory cement of a compo corresponding to a mixt of fireclay, SiO_2 and asbestos fiber is often used. Fireclay bricks are generally satisfactory. Silica bricks are used for the roof and the parts of the sides not coming in contact with slag. Pure magnesia bricks, fully burned and enclosed in steel cases, show long life. Zirconia bricks are highly refractory but very expensive. The use of fritted linings has increased widely. The range of metals for the linings of industrial furnaces of the Ajax Wyatt type and

has increased greatly the no of runs obtainable without relining the furnace

W. H. BOYNTON

Refractory clays in Calhoun and Pike counties, Illinois. J E LAMAR, *Ill State Geol Survey, Rept Investigations No 22*, 43 pp (1931)

R. F. FERGUSON

The determination of aluminum oxide in refractory clays. R. BIAZZO, *Ann. chim applicata* 21, 3-12 (1931)—Where only the amt of Al_2O_3 is to be detd, time is saved by volatilizing the SiO_2 by means of H_2SO_4 , HCl , detg $Al_2O_3 + Fe_2O_3$ and then detg Fe_2O_3 in the ordinary ways and obtaining the amt of Al_2O_3 present by difference Al_2O_3 was detd with less than 0.25% error by this method

A. W. CONTIERI

The possibility of andalusite as a refractory. FRANK H. RIDDLE, *Trans Electrochem Soc* 59, preprint 16 pp (1931)—The com development of $Al_2O_3-SiO_2$ refractories stimulated the search for deposits of a natural mineral of this so-called 'sillimanite type'. Regardless of the proportions in which the 2 oxides occur to begin with, after exposure to furnace temps. mullite, $3Al_2O_3 \cdot 2SiO_2$, is formed with one or the other oxide in excess. Many expts have shown that andalusite of proper grain size and proportion, when held together with a min amt of a plastic aluminous bond such as clay, makes a very desirable refractory. It is const in vol during calcining. An extensive deposit of andalusite has been located in Mono County, Calif. About 350 million spark plug cores have been made from this andalusite. Important localities of cyanite, sillimanite and dumortierite are briefly described. Tests on andalusite refractory are detailed, demonstrating its superiority over other refractories of this class.

C. G. F.

Works tests on refractories and service conditions. R. J. SARJANT, *Trans Ceram Soc (England)* 30, 46-65 (1931)—Load tests indicate serviceability. Changes in porosity and d on firing to 1450° give useful indication as to probable service behavior. Simulative slag resistance tests are valuable if classified according to type of refractory and use. Reliable spalling tests are hard to get and to interpret in terms of service conditions. Heat cond. data have definite economic value. For magnesite refractories, optical detn of the % periclase present and the d and porosity changes before and after firing to 1600° are important tests. Chem analysis, particularly as regards SiO_2 , MgO and CaO , provides the most essential information regarding dolomite refractories. In silica refractories the degree of conversion and bond are most important characteristics, while in fireclay bodies bonding of the clay is very important. Refractory cements should be examd by all available tests both alone and built up with refractory units.

H. F. KRIEGER

Thermotechnical possibilities and requirements in the ceramic, especially in the refractory-materials industry. W. BRASS, *Sprechsaal* 62, 445-7 (1929)—Thermotechnical conditions in the ceramic industry and their causes are investigated. Good results were obtained in tests carried out on an old chamber oven making use of gas-analytical app.

E. I. S.

The effect of smelter atmospheres on the quality of enamels for sheet steel. ANDREW I. ANDREWS and EMANUEL A. HERTZELL, *Univ Ill Eng Expt Sta, Bull No 224*, 15 pp (1931), cl C A 25, 1650—The presence of N_2 , CO_2 and reducing atm in the smelting furnace is not deleterious to sheet steel enamels, but SO_2 is extremely harmful.

H. L. OLIN

Notes on the manufacture in New Zealand of electrical porcelains. E. W. ACKLAND, *Proc. New Zealand Soc Civil Eng* 16, 274-86, 287-302 (1929-30)—This is a 98% local industry. Methods of manu are outlined.

E. I. S.

A study of the electrical strength of porcelain. JOHN O. KRAHENBUHL and CULLEN W. PARMELFE, *Trans Electrochem Soc* 59, preprint 15 pp (1931)—The paper records some of the existing inconsistencies in the literature concerning solid dielectrics, particularly porcelain. Observations show that the puncture track is largely a matter of chance and depends upon flaws. A statistical study was made of a particular body, prepd under controlled conditions. The paper does not try to define a test method or a specification, but shows the importance of a complete study of the test data.

C. G. F.

The microstructure of some porcelain glazes. CLYDE L. THOMPSON, *Univ Ill Eng Expt. Sta, Bull* 225, 24 pp (1931)—Glazes studied fell within the compn range of (0.3 K_2O , 0.7 CaO), 0.4 to 1.0 Al_2O_3 , 2.0 to 8.0 SiO_2 . A series of glazes was made from the usual potters' materials and examd microscopically. A second series was made from pure chemicals and the primary crystalline phase sepg from each glass and its temp of equilibrium were detd. Conclusions (1) Good mat glazes owe their matness to crystn. (2) Al_2O_3 plays a double role in mat glazes, furnishing an essential constituent of the crystallizing mineral and also regulating the viscosity to prevent

under crystal growth. (3) Many of the bright patches in glasses are not all glass but contain undissolved quartz, and show traces of devitrification. (4) An excess of either Al_2O_3 or SiO_2 will cause devitrification, while a greater excess will cause immaturity. (5) The Na_2O introduced by the feldspar has a definite effect on the composition of the mineral sepi from the melt upon crystals. H. L. OLRY

Master-surface leopards (EAMES) 24. Prevention of explosions in mirror-silvering materials (EAMES) 24.

Glass. CORNING GLASS WORKS. Fr 677,570, June 17, 1930. Glass which has been frosted e. g., elec. bulbs, is prevented from "aging" by washing it in an aq. soln. of AlF_3 .

Glass making. ERNST HINGERS. Fr 677,418, June 14, 1930. App. for rolling plates of glass is described.

Melting glass. GUSTAV KEFFELER. Fr 677,544, June 24, 1930. In melting glass the melt, contg. finely divided SiO_2 , sweated in known manner, is subjected to a brief elevation of temp. so that leaping of the foam has a more gradual and the sweating material, before melting, remains as completely as possible in the action of the flame.

Paddle mechanism for circulating molten glass. JONES MOWATRY, LTD., and ALEXANDER F. McLENNAN. Ger 522,549, Feb. 15, 1929. Brit. 220,153 (C. A. 24, 276).

Apparatus for feeding molten glass. THOMAS F. FRANKSON. Ger 522,550, April 25, 1929. Corresponds to Brit. 221,173.

Mold for glassware. WM. T. BAKER, JR. (to Hartford-Empire Co.). U. S. 1,778,126, March 31. Structural features of a mold with pyres in its walls for circulation of a temp.-controlling fluid.

Glass-forming apparatus suitable for manufacture of blow-molded ware. EDWARD H. LEHNEZ (to Hartford-Empire Co.). U. S. 1,778,191, April 7. Structural features.

Mold for forming glass articles such as bottles. CLAUDE C. MORAN (to Tygart Valley Glass Co.). U. S. 1,778,217, March 31. Mech. features.

Apparatus for manufacture of stemmed and footed glassware. VIRGIN O. CORNWELL (to Federal Glass Co.). U. S. 1,778,278, March 31. Structural features.

Decorating and annealing leet for glassware. VIRGIN M. MCKEAND (to Hartford-Empire Co.). U. S. 1,778,322, March 31. Structural features.

Electric furnace for glass manufacture. HARRY F. HUNTER (to Pittsburgh Plate Glass Co.). U. S. 1,779,271, April 7. Structural and elec. features.

Cooling furnace for sheets of glass. N. V. M. J. TOE BEHRENS EN EXPLOITATIE VAN OCTROOIEN. Fr 677,572, June 24, 1930.

Roll-conveyor leet for intermittently prepared sheet glass. N. V. M. J. TOE BEHRENS EN EXPLOITATIE VAN OCTROOIEN. Ger 522,573, June 26, 1929.

Roller apparatus for conveying intermittently prepared sheet glass to the leet. S. C. ARON (to MANUFACTURERS OF GLASS BY PRODUCTS) ORIGINES DE ST. GOBAIN, CHATEL & CHRY. Ger 522,416, Oct. 14, 1927.

Method and plant for making sheet glass with wire insertions. HERMANN MÜLLER-SCHNEIDER. Ger 522,587, July 17, 1928.

Sheet- and glass-glass manufacture. WM. WILHELMUTH (to Libbey-Owens-Ford Glass Co.). U. S. 1,779,223, April 7. Mech. features.

Method of making glass articles showing marbled or veined effects. JEROME GLASSWARE SCOTT & COY. Ger 522,417, Dec. 4, 1928.

Laminated glass. GEORGE B. WATKINS (to Libbey-Owens-Ford Glass Co.). U. S. 1,779,670, March 31. Laminated products such as those comprising glass sheets with an intervening sheet of cellulose or the like are subjected to the action of a solvent bath such as tracetin for a sufficient time to cause a "flowing" of the intervening sheet material to form a seal at the edges. Cf. C. A. 24, 276.

Kiln for firing ceramic ware. ROBERT H. YOUNGMAN and ROBERT H. H. PIERCE. U. S. 1,779,297, March 31.

Kiln (with preheating, firing and cooling zones) for treating ceramic ware. FRANK M. HARTMAN (to Harry Ceramic Service Co.). U. S. 1,779,569, April 7.

Direct-fired tunnel kiln and heating system for ceramic articles. ISAAC HARTER, ARTHUR M. KOPPEL and FREDERICK H. NORTON (to Babcock & Wilcox Co.). U. S. 1,779,296, April 7. Various details of construction and operation are described.

Channel ceramic furnace. F. A. MERRICK. Rem. 19,510, Feb. 28, 1931.

Brick kiln. BOHRERSSON & CO. A. S. H. Ger 516,970, Feb. 13, 1929. Details of fuel loading are given.

Insulating block. EDWIN B. FORSE (to Carborundum Co.). U. S. 1,726,934.

March 31. An insulating brick suitable for lining furnaces is made with a dense facing, having a permeability of less than 100 units through the brick and facing and a permeability of over 300 units when the facing is removed. The facing may be formed of vitrified clay 4, unbaked clay 4 and Na silicate 2 parts.

Refractory material. SANDROD S. COLE (to Koppers Co.) U. S. 1,798,072, March 31. A refractory material suitable for use in furnace or coke-oven construction is formed from a silica body such as ganister mixed with a ferric salt such as FeCl_3 , lime, an org. acid such as lactic acid and with $\text{Ca}(\text{CNS})_2$ and water, molded, dried and fired. Cf. C. A. 24, 3973.

Abrasives. BAKELITE CORP. Fr. 698,090, June 26, 1930. A supple phenol resin such as that described in U. S. 1,590,079 (Byck, C. A. 20, 3090) is used for binding abrasives to supple supports.

Enameling iron. VEREINIGTE CHEM. FAB. KREIOL, HELLER & CO. Ger. 522,472, Nov. 4, 1923. Fe articles, particularly sheet Fe, are provided with a colorless ground-enamel coating, free from CoO or NiO by applying a suitable mixt. of readily fusible and difficultly fusible substances in the form of a moist pulp and then heating the article for a short time, so that the difficultly fusible substances do not completely dissolve in the readily fusible substances with the result that a coarse-grained coating is produced. A suitable mixt. contains borax 36.3, feldspar 36.3, quartz 32, soda 6, NaNO_3 6, CaF_2 1.8 and kaolin 10 parts.

20—CEMENT AND OTHER BUILDING MATERIALS

J. C. WITT

Chemical constitution of cements. FERRARI. *Ind. cemento* 25, 6(1028); *Ceram. Abstracts* 9, 259.—F. reviews the literature dealing with the chem. constitution of calcareous cements and cites references dating between 1785 and 1925. G. G.

White cements. L. NICOL. *Baumarkt* 29, 451-2(1930); *Ceram. Abstracts* 9, 1017-8.—White cements are classified as portland, "marble," magnesia and special cements. White portland cements include all those which chemically are similar in compn. and in properties to the gray portland cement. The white portland cements examd. had on the av. the following compn.: lime 51-2, SiO_2 13-24, Al_2O_3 6-13, Fe_2O_3 0.2-0.6, SO_3 0.4-1.5, MgO 0.2-0.4, potash salt 0.3-2.0, loss on ignition 2.3-6.0%. In every white portland cement the Fe_2O_3 was replaced by Al_2O_3 down to a small part; its properties are similar to those of the ordinary gray portland cement, except that the white cement sets more slowly because of the higher degree of gel formation. From 10 to 25% of fine white quartz sand should be added to the white cement in making the concrete, and the mixing water must contain no Fe. A plant at Stettin made a good white Stern cement of the following av. compn.: lime 58-60, SiO_2 19-20, Al_2O_3 12-14, Fe_2O_3 0.3-0.6, SO_3 1.0-1.5, loss on ignition 5-6, MgO 0.1-0.3, and potash soda 0.3-0.5%. A good French white cement is the Le Noble cement, which has the following compn.: lime 62.75, SiO_2 22.00, Al_2O_3 5.90, Fe_2O_3 0.25, MgO 0.60, SO_3 0.55, alkalis 0.20, loss on ignition 9.15%. "Marble" cements are, in general, not to be used for outside work, for they are not hydraulically hardening cements. They comprise double-calcined gypsum with adds. of alum. A Luxemburgian marble cement has the following compn.: lime 67.80, SiO_2 20.20, Al_2O_3 1.25, Fe_2O_3 0.25, MgO 1.38, CO_2 1.26, "sand" 1.30, loss on ignition 6.04, and SO_3 0.40%. Russia has a very light marble cement called La Glyptolithe. A Swiss white marble cement has the following compn.: ignition loss 5, lime 39, SiO_2 18, Al_2O_3 and Fe_2O_3 23, SO_3 7, alkalis 1.2, MgO 0.6, and barite 0.2%. Of the white special cements, one is made from rock chalk, infusorial earth (kieselguhr) and soda soln., and the other from calcined lime (white lime) and ground colorless silicic acid which are mixed in the slaking drum. G. G.

Action of magnesium sulfate on cement. P. KREMSER. *Ind. cemento* 26, 8 (1929); *Ceram. Abstracts* 9, 258.—K. describes the occurrence of MgSO_4 and its chem. relation to cement. CaO is subject to the disintegrating action of MgSO_4 according to the equation $\text{Ca}(\text{OH})_2 + \text{MgSO}_4 \rightarrow \text{CaSO}_4 + \text{Mg}(\text{OH})_2$; the hydrate of magnesia being a white, gelatinous mass. The detrimental compd. sulfatoaluminate of Ca is then formed according to the equation $3\text{CaSO}_4 + (3\text{CaO} \cdot \text{Al}_2\text{O}_3 \cdot 12\text{H}_2\text{O}) + 18\text{H}_2\text{O} \rightarrow 3\text{CaO} \cdot \text{Al}_2\text{O}_3 \cdot 3\text{CaSO}_4 \cdot 30\text{H}_2\text{O}$. This action effects change of vol. and disintegration of the cement. K. also describes results obtained by satg. the specimens with "Sika," a cement paint for reducing disintegration by MgSO_4 . G. G.

Material balance in the rotary-kiln cement plant. D. STEINER. *Tonind-Ztg.*

55, 470-80(1931)—A 100-kg charge yields 64.5 kg of clinker and 5 kg of blue dust. This latter is composed of 1 kg unburnt charge, 2 kg calcined, 1.3 kg ash and 0.7 kg coke. With 330 g of coal per kg of clinker and 15% excess air, 3.35 kg of air per kg of clinker is required. A flow sheet and analytical data are given. D. A. SOULZ.

Specifications for the grading of sand, gravel and crushed aggregates. ANON. *Zement* 20, 37-8(1931)—A report is given of the 3rd Conference of the Comm. for the Standardization of Particle Sizes. H. F. K.

Chemical studies of rock used in the construction of old monuments in the city of Salamanca, of the patina of the same, and its artificial reproduction. I. RIBAS AND PETRA DE PRADA. *Anales soc. españ. fis. quim.* 29, 211-20(1931)—The golden yellow color of the patina on the old stones in Salamanca is due to the hydrosulfide. Patination is caused by dehydration of the hydrosulfide on the surface. Suitable treatment of the stone with pptd. $\text{Fe}(\text{OH})_3$ reproduces the patina. E. M. SYMMES.

Pulverized iron for waterproofing and hardening floors and walls. CHARLES Kew. *Creamery and Milk Plant Monthly* 20, No. 6, 90-1(1931)—Pulverized iron suspended in water is applied to concrete, brick or masonry and permanently waterproofs these materials. The water, soaking the walls, carries fine particles of iron with it which oxidize and swell up and thus seal the pores. A. H. JOHNSON.

Colonial wood from Spanish Guiana. H. T. BAYTECAS. *Anales soc. españ. fis. quim.* 29, 184-8(1931), cf. C. A. 25, 189—Detrs. of water, ash, B. t. u., apparent d., and ultimate analysis were made on *Dyospiros exilis* (ebony), *Pterocarpus sovanusii* (palo rojo or padouk), *Lourea klaineana* (civero), *Isolana klaineana* (ogoa or n'f6), ebanan or m. bebaime (a species of *Chrysophyllum*), and adela or ondondu. E. M. SYMMES.

Acid resisting materials for chemical and building industries (GREGOR'EV, SILVESTROVICH) 13. Kinetics of transformation of the various forms and stages of hydration of Ca sulfate (ГЕРГЕНЦИТ) 2. Cleaning [rotary cement kiln gases] electrically (Ger. pat. 522,389) 4.

WALE, KURT. Die heutigen Erkenntnisse über die Wasserdurchlässigkeit des Mörtels und des Betons. Berlin. Ernst 02 pp. M 9.

Cement. G. POLYSIUS A-G. Ger. 522,022, Dec. 4, 1923. A method of de-watering crude cement sludge in a horizontal centrifuge is described.

Cement. BRUNO BRUNN. Fr. 697,673, June 19, 1930. The relative content of SiO_2 in normal portland cement is reduced by increasing the content of sesquioxides and the mass is fused. In the fusion of a mixt. of slag and lime the relative content of SiO_2 is reduced by increasing the content of sesquioxides. Other fluxes such as CaCl_2 or compds. of HF may also be added.

Cement. CASPARD WINKLER. Fr. 698,422, July 5, 1930. Cements of various kinds are protected against the action of corrosive waters by replacing a part of the sand by clinker of portland cement or other cements, the size of grain being between 0.15 and 2 mm. The clinker is preferably treated before use with solns. of alkali or alk. earth compds. or F. compds. or solns. of oxalic acid or its salts.

Cement. TREPIERRES ET LAMINOIRS DU HAVRE (ANCIENS ÉTABLISSEMENTS LAZARE WEILLER) and SOC. COOPÉRATIVE DE RUGLES ET LA CANALISATION ÉLECTRIQUE RÉUNIS. Fr. 698,515, May 24, 1930. In making porous cements by decomposition of part of the mixing water under the action of Al powder, the formation of a protecting layer of Al_2O_3 on the Al is prevented by the addition of certain chlorides such as MgCl_2 or HCl and calcined MgO . A soln. of KMnO_4 may also be used to prevent oxidation.

Cement. ARTHUR F. FARBE. Fr. 698,520, May 6, 1930. A special portland cement is made by substituting pozzolana for CaSO_4 in com. portland cement.

Fused cement. STEFAN KOHUT. Ger. 516,839, Nov. 23, 1928. See Fr. 685,027 (C. A. 24, 5456).

Porous cement. SOC. ANON. TREPIERRES ET LAMINOIRS DU HAVRE (ANCIENS ÉTABLISSEMENTS LAZARE WEILLER) and SOC. COOPÉRATIVE DE RUGLES ET LA CANALISATION ÉLECTRIQUE RÉUNIS. Fr. 698,320, July 3, 1930. Porous cements are made by the action of a metal such as Al and a soln. of KMnO_4 or MgCl_2 or MnCl_2 , CaO being added in suitable proportion to increase the action of the metal.

Dry mixing of cement-forming materials. MIKHAEL VOGEL-JORGENSEN (to F. L. Smidth & Co.). U. S. 1,798,423, March 31. Various details of procedure and app. are described for obtaining a uniform mixt.

Waterproofing material for cement. MANZIRO NAKAHARA and TOKYO INDUSTRIAL

LABORATORY Japan. 89,619, Dec. 19, 1930 A mixt. of resin and pitch is emulsified with aq soln of glue, and H_2SiF_6 and a substance contg sol silicic acid are added.

Concrete. **BETONIT GRS** M B H and N V. HOLLANDSCHE BETONIT MAAT-SCHAPPIJ 1r 698,636, July 8, 1930 Articles such as pipes molded in concrete are placed in a closed space, which is then filled with liquid bitumen, asphalt or tar and submitted to pressure

Preparing and placing concrete. **KARL P BILNFR** (to Aerocrete Corp. of America) U S 1,798,924, March 31 An intimate mixt is formed of cement, water and a gas-producing agent such as Al powder, and a counteracting pressure is applied to the mixt to prevent expansion of gas bubbles and to permit a pressure to develop within the mass while the chem reaction is in its incipient stages, the mass is then transported to the place of use, the counteracting pressure is relieved and it is deposited where desired App is described

Slate as a concrete aggregate. **P N SHADLIVSKY** and **M I KOIFMAN** Russ 19,516, Feb 28, 1931 Crushed slate, obtained as a waste product in slate production, is heated at 1100–1200° before being incorporated into concrete

Coating concrete vats. **STANDPASSWERB ROSTOCK & BAFRLOCHER** and **GEORG WALTER** 1r 698,562, June 25, 1930 Fermentation vats and reservoirs are coated internally by aq emulsions of bitumens, resins or lignite waxes to which oils and fats may be added Mortar may be added to increase the adherence

Road-making compositions. **ROMAN GRANGG** and **ANNEMARIE SCHMÖLZER** Austrian 121,773, Oct. 15, 1930 The binding of compns comprising rubble or the like and water glass, is improved by addn of powd sintered magnesite, or of powd wood charcoal that has been stored in CO_2 The sintered ferruginous waste obtained in working up magnesite may also be used

Road tar. **THERMAL INDUSTRIAL & CHEMICAL (T I. C.) RESEARCH CO., LTD** Fr. 897,824, June 23, 1930 A tar suitable for roads is made by adding a small proportion of acid, preferably H_2SO_4 , to a crude tar free from wax, agitating the mixt at 55–83°, neutralizing the mixt and dehydrating it, and if necessary heating it to volatilize the light oils and obtain the desired consistency

Asphaltic and bituminous compositions. **GENERAL TECHNICAL CO., LTD** Fr 698,686, Oct. 7, 1929 Compns. of relatively high m p suitable for roads are made by heating residues from the distn of natural petroleum or coal tars to a temp of 360–430° at atm pressure, and eliminating the resulting liquid products from this partial disson as they are formed

Expansion joint for concrete pavements. **ALBERT C. FISCHER** (to Philip Carey Mfg Co.). U S 1,800,193, April 7. Various details are described of the use of fibrous material and bituminous material together. U S 1,800,194 relates to products such as satd. paper or felt having an adhesive surface such as oil or asphaltic compn protected by a water-sol covering such as Na silicate. Cl C A 24, 4608

Building block. **FRANCIS J STRAUB** U S 1,794,505, March 31. The smaller particles and fines are sepd from an original mass of cellular coal-cinder aggregate and while dry these are mixed with a suitable binder such as portland cement. The remainder, composed of the larger pieces, is soaked in water and the materials are then mixed together for molding into blocks. An arrangement of app is described

Shaft furnace for burning gypsum. **I. F FLORETZOV** Russ 19,520, Feb 28, 1931. Structural details.

Artificial stones, etc. **SOC ANON POUR L'IND DE LA MAGNÉSITE (SUCCURSALE DE HONGRIE).** Fr 897,898, June 25, 1930 Very refractory stones resistant to rapid changes of temp are made by grinding magnesite (fritted or conereted) or overburnt MgO , removing the very fine particles—not more than 50%—and replacing them by an equal amt of finely ground substances rich in Al_2O_3 . The mixt is used as mortar, fireclay lining or molded in the form of stones and baked.

Fireproof composition suitable for use in building construction. **CARL MARCZINEK.** U. S. 1,800,024, April 7 Volcanic cinders 45, cement 20, gypsum 5, hydrated lime 5 and finely ground pumice 25 parts are mixed with sufficient water to cause the mixt. to set.

Apparatus for making asbestos-cement tubes. **RUDOLF HAYDEN** Austrian 121,643, Oct 15, 1930.

Surface coloring of flat articles such as asbestos-cement shingles. **CLEMENTS BATCHELLER** (to Bemis Industries, Inc) U. S 1,798,990, March 31. Flat objects having an alk. reaction are treated with a soln of a metallic salt such as Cu and Fe salts which will react with the alk substance to form an insol. compd and treated surfaces are juxtaposed to retard evapn. of the soln.

Surface coloring of articles such as asbestos-cement shingles. CLEMENTS BACHELLER (to Bemis Industries, Inc.) U. S. 1,708,997, March 31. In treating an article having an alk. reaction and having inorg. coloring matter mechanically incorporated in its surface, to bind the coloring matter in the surface, the surface is treated with a soln. of a metallic salt. U. S. 1,708,998 describes the similar surface coloration of shingles comprising cement and fibrous material. U. S. 1,708,999 relates to coloring porous bodies such as asbestos-cement shingles by applying a soln. of a metal salt such as FeSO_4 or CoSO_4 , which is capable of oxidation and then applying a reacting substance such as a peroxide to form insol. coloring material in the surface portion of the material. Various details and examples are given.

Plaster mixture. GEORGE H. A. RUBY. U. S. 1,799,324, April 7. A material capable of holding nails firmly without splitting or spalling is formed of cotton fibers with a plaster-of-Paris binder solidified through crystal after addn. of water. U. S. 1,799,325 also relates to similar material.

Plaster-board. THEODORE E. KNOWLTON (to Certain-Teed Products Corp.) U. S. 1,798,603, March 31. Mech. features.

Drying wood, etc. ARTIEROLAGET BANKFIRMAN ALFRED BERG. Ger. 522,141, Sept. 24, 1929. The materials are treated in a closed container with warm air or other gas under pressure and the pressure is then reduced to zero as rapidly as possible. The treatment may be repeated. App. is described.

Artificially drying wood. FRIEDRICH MOLL. Ger. 516,064, Mar. 20, 1930. The drying chamber has an opening from which observations and samples can be taken during the process.

Preserving wood. GRUBENHOLZIMPRAGNERUNG G M B H. Fr. 698,624, July 7, 1930. Wood is preserved by a mixt. of equal parts of K_2CrO_7 and NaF .

21—FUELS, GAS, TAR AND COKE

A. C. FIELDNER AND ALFRED H. EMERY

Investigation of the aluminum apparatus for the testing of spontaneous combustion of anthracite coal and other materials. D. J. W. KAZULPH. *Brennstoff-Chem.* 12, 107-11 (1931); cf. C. A. 25, 394. —By use of app. previously described (C. A. 25, 394) the influence of heating and O_2 flow rates, sample and particle size upon the results obtained has been detd. The best results are obtained when CO_2 is passed through the coal at 6 l./hr. until the app. has reached 120° . When the coal has reached 118° , O_2 is started through at 10 l./hr. and the temp. of the app. raised to the kindling point of the coal. Depending upon the shape of the plotted temp. curves of the app. and coal at the kindling temp., rates of 1° or 2° or $3^\circ/\text{min}$ are used. For each detn. 5 cc. of briquetted coal sized between 10 and 20 mesh/cm. is used. Anthracite, brown coal and wood charcoal have been used in showing graphically the effects of the variations. F. W. J.

New method for practical fuel examination and calculation. M. DOLCH. *Brennstoff und Wärmewiss.* 12, 253-67 (1930). —Review of the author's previous work (cf. C. A. 24, 221, 5969).

Ignition and combustion of fuels of low heating value. MARCAND. *Die Wärme* 54, 208-13 (1931). —Crit. data concerning the ignition characteristics of fuels are tabulated and discussed. The influence of the construction of the boiler on the ignition and combustion processes, and the importance of catalysis on ignition of fuels are discussed.

ALLEN S. SMITH

Determination of the heating value by a simplified calorimeter. A. P. SHAKIRNO AND M. D. ZHUKOVSKI. *Izvestiya Teplotekhn. Inst. (Trans. Thermo-Tech. Inst.)* 1931, No. 1, 34-42. —The heating values of various coals and wood were detd. in different calorimeters. The Parr calorimeter is recommended provided that the Na_2O contains not less than 92% of active O_2 and the proper amounts of Na_2O and tartaric acid are used, without the addn. of $\text{K}_2\text{S}_2\text{O}_8$. Corrections for radiation must be made, and max. deviations of 1.16% were observed in various detns.

A. A. BOESHLINCK

Wood as a fuel for internal-combustion engines. G. DUPONT. *Bull. inst. pan.* 1931, 5-8, 31-6, 65-71. —An address describing the work carried out recently by D. and co-workers at the Institut du Pm (cf. C. A. 23, 2551, 24, 1490, 4914).

A. PAINTEAU COUTURE

Conversion tables for recalculating to various laboratory conditions results of fuel analysis and ash determinations. V. ALEXANDROVSKI. *Izvestiya Teplotekhn. Inst. (Trans. Thermo-Tech. Inst.)* 1931, No. 1, 56-60. —Tables are given for recalculating the results of lab. analysis to the as-received basis and to the dry basis.

A. A. BOESHLINCK

Determination of the ash content of coke and coal. H. A. J. PIETERS. *Helv. Chim. Acta* 51, 77-9 (1931). —From comparative ash detns on coke and coal ground in a hammer mill, rolled on a steel plate or pulverized in a mortar it is shown that Fe particles can cause variations in ash content up to 0.4% (highest error for rolled coke). A particularly hard material needs to be used for the construction of grinders. A temp of 900° is advocated for ash detns., higher temps cause losses. Cooling of the samples for moisture or ash detns. in air is equally as good as in a desiccator, as long as it does not take more than 10 min. Tables of check analyses are given. B. J. C. VAN DER H. H.

Fuel technology and the classification of coal. CLARENCE A. SEYLER. *Proc. S. Wales Inst. Eng.* 47, 6-9 (1931). —The properties of coal are divided into 2 categories: the conen. characteristics and the mobility or reactivity characteristics. The conen. characteristics are those which depend on the conen. of the combustible elements or the energy in the coal or its products of combustion. They include the emmpn. of the coal, its calorific value, vol. of air required for combustion, vol. and emmpn. of flue gases, flame temp. of gases and combustion efficiency. The relation between the vol. of flue gases or air and the net calorific value of the coal is a special case of a general relation which can be deduced from Dulong's law. The plotted lines of isocals therefore represent all other conen. characteristics. The mobility or reactivity characteristics depend on the freedom with which the colloidal particles or those of mol. dimensions can move and undergo rearrangement or enter into combination with external reagents. Coking properties belong in this category. Softening points detd. by the Charpy and Durand method show great regularity when plotted and lines of equal softening points may be drawn. The lowest line surrounds the best coking coals, the higher the line the poorer the coking properties. ALDEN H. EMERY.

Petrography and coal classification. CLARENCE A. SEYLER. *Proc. S. Wales Inst. Eng.* 47, 4-6 (1931). —S. briefly reviews 3 systems for the classification of coal lithological (vitrain, clarain, durain and fusain), botanical (anthraxylon and attritus), and micropetrographical (humic material, resins, resistant plant remains and fusain). The latter are in the main proximate chem. components. ALDEN H. EMERY.

The caking power according to Meurice as a basis for classification of coal. D. STAVORINUS. *Helv. Chim. Acta* 51, 79-82 (1931), cf. M. C. A. 8, 1496. —The Meurice method of detg. caking power is reviewed. The "limiting agglutination index" is detd. by the no. of g. of sand which 1 g. of coal can bind without giving more than 1 g. loose powder. The sand passing a no. 40, and retained on a no. 50, sieve is heated with coal ground to pass a no. 80 sieve in a 25-cc. covered Messner porcelain crucible, between 900 and 950° in a muffle until the flame between crucible and cover has disappeared. The loose powder remaining is weighed. A second significant figure, the "cokefaction index," is detd. from the compressive strength K of the cake of 17 g. sand + 1 g. coal. If D is the wt. of the loose powder after caking, the index is $17K/D$. The usual (European) classification of coal corresponds to limiting indexes as follows: flame coal (volatile matter more than 35%) 0 to 8, fat gas coal (volatile > 23%) at least 10, coking coal (volatile 18 to 32%) above 14, various fat coals (volatile 11 to 22%) from >10 to >2, lean coal (volatile 9 to 11%) 2, anthracite (volatile up to 8%) index 0. For 15 samples of Ruhr and Durham gas coals the cokefaction index is given. It varies for Ruhr coal from 40 to 425, K between 3.5 and 14, D from 0.30 to 1.5, for Durham coal the index is 196 to 1135, K from 5 to 10, D 0.30 to 0.52. For the detn. of compressive strength the Lechman app. was used. B. J. C. VAN DER HOEVEN.

A note on the mineralogy of coal as suggested by x-ray examination. HENRY BRIGGS. *Colliery Guardian* 138, 638-40 (1929). —Mineral impurities in coal are classified under four heads: (1) those existing in the coal substance, and so finely disseminated through it that they cannot be extd. by any known phys. process, (2) those forming layers, streaks or flakes lying in the bedding planes, (3) those forming streaks, flats, flakes or plates, usually at angles approaching the right angle, and (4) particles, crystals or irregular masses, large or small, scattered, apparently promiscuously through the coal. The first is discernible only in the radiograph of a coal band as a uniform and a featureless cloudiness of the clearest portions, while the other 3 varieties are sharply silhouetted in an x-ray photograph, and their manner of occurrence is still more clearly presented by the stereoscopic method of Kemp and Thomson. A review is made of characteristics of the mineral impurities, and the low ash content (compared to that estimated from the plant source of the coal) is explained. X-rays provide refined methods of qual. analysis not previously available and at the same time enable the manner of distribution of these foreign substances to be studied without destroying the sample. W. H. BOYNTON.

Coal properties and their applications. T. KIDOKORO AND G. SHIRANE. *J. Fuel*

See JONES 9, 43-6 (1929) —A method is described for the identification of the geological rank of coals, lignites, etc., by their treatment with oxidizing agents followed by pyrolysis of the benzoic acids with NaOH. Expts. on the evolution and spontaneous evolution of coals, with special reference to their rank as dealt above, are also discussed. B. C. A.

The preparation of coal dust for and its analysis by projected microscopic images. EATON STACE. *Fuel* 12, 147-50 (1931), cf. C. A. 23, 5229 —The procedure and app. are described in detail for the examn. of coal particles, reduced to 0.25-0.88 mm. size. For mounting, a material prepd. from 3 parts diamond, 2 parts shellac and 1 part Venetian turpentine or either montan or ceramium wax alone is used. To 8 parts by vol. of this medium, 6 parts of the coal is added. The mixt. is then poured into a 2 x 2 x 1 cm. mold. After the setting and removal, progressively finer grinding and wet polishing with the final use of alumina to produce relief, a surface of the block is prepd. for microscopic examn. at 15-20 diam. magnification. F. W. JONES

Notes on occurrence of spontaneous heating in coal mines. E. H. VALENT. *Colliers' Gazette* 141, 1971-2, (1931) —For the Seven Foot Sashbury seam at the Birchwood Colliery it appears that the CO₂/O₂ absorbed ratio of 1% is a critical figure. Having reached this point, it may recede or after the ratio has remained stationary for several days, heating may develop with startling rapidity. It was found that the CO₂/O₂ absorbed ratio may increase from 1.1 to 2.9% in less than 10 hrs. and to a stage when the 1st change in smell is apparent in 3 hrs. V. emphasizes the importance of analyses every 24 hrs. in addition to the frequent examn. of the working districts of such mines for any change in the smell of the return air. Such analyses should include air in the working places and the airways near the face to discover the development of heating in its early stages. H. C. PARKER

Process of combustion of pulverized coal. A. CASATI. *Compt. rend.* 192, 567-9 (1931) —Pulverized and dry (1% H₂O) coals (lignite or anthracite) ignite at characteristic temps. between 70° and 90°, temps. which correspond to the max. disengagement of H₂ during dry fractional distn. This leads to the conception that ignition of coal starts with disintegration of the H₂ disengaged, the carbonization continues along with the progressive and exothermic evolution by air, first of the H₂, then of the transformed coals, and then of CO. The formed monolayer coke reacts with the H₂O present and the resultant CO + H₂ burns. This conception differs from the intermediate formation of the water-gas theory of C. A. H. (Brennstoff und Verhüttung, Berlin, 1921), and is supported by recent analyses of Fick (Le Châtelier des Chaudières au Charbon Pulvérisé, Paris, 1930, p. 57), the results of which are pointed to show that C is first oxidized to CO rather than to CO₂. ALBERT THOMAS FELLOWS

Graphic estimation of the combustion process. ALBERT THOMAS FELLOWS. *Dr. Wayne* 14, 221-4, (1931) —A method of regulation, and estm. of combustion efficiency is described which eliminates the necessity of flame analyses and tedious calcs., and which is simple and reliable. Oswald line-gas diagrams (C. A. 14, 257) are prepd. from equations representing complete and incomplete combustion of gaseous, liquid, and solid fuels. From these diagrams the characteristics of the combustion may be detd. The influence of CH₄ and H on the accuracy of the diagrams is considered. ALLEN S. SMITH

The specific heats of gases involved in combustion. A. JAROSCH. *Gastech* 59, 947 (1929), *J. Soc. Gas Tech.* 14, 144 —The sp. heats of individual gases, as well as those of typical mixts. of waste gases, are shown in tables and diagrams. The math. basis is given together with examples of the detn. of (1) the heat dissipated by gases while cooling and continuously discharging in pressure, (2) the mean sp. heat of a gas between two specified temps., (3) the heat lost in the waste gases. G. G.

Saskatchewan Lignite. J. P. SAWYER. *Can. Mining J.* 52, 420-5, (1931), cf. C. A. 25, 199 —Use is being materially extended within the province by adoption of suitable grates and stokers. W. H. BIRYON

The gas plant at Ivey. POTTERMAN. *J. Amer. gas* 55, 69-74, 94-102, (1931). —Historical. Operation details are given with many drawings. B. J. C. v. n. H.

The modernization of the gas plant at Bremer. G. DEBAN. *J. Amer. gas* 55, 25-7 (1931). —A short description of the new installation of 7 vertical discontinuous chamber ovens, each of a capacity of 1500 kg. coal in 12 hrs. Steam is injected during the last 2 hrs. of carbonization. The plant has built-in producers. B. J. C. v. n. H.

Gas production and its economics in Germany. SIEBANDER. *Gesellschafts Nachrichten* 1, No. 1, 25-6, (1930) —An historical review of the gas production, its application for domestic and other purposes and the potential capacity of German oil fields are given. Statistical data on gasifier recovery from wet gases are tabulated. A. A. B.

Plant for preparation of coal and coke at the gas plant in Niederwül. J.

ANDREGO Schweis l'er. Gas- u Wasserfach Monats Bull 11, 78-82(1931)—A short description with drawings. B. J. C. VAN DER HORVEN

Fundamental aspects of the natural-gas situation in Chicago. SAMUEL INSULL, JR. *J. Western Soc Eng* 36, 111-5(1931)—A brief résumé of the basic economics which shows that only those domestic consumers using large amts. of gas will benefit greatly by the use of natural gas. W. H. BOYNTON

Heat loss of gas furnaces. FARRZ SCHUSTER *Die Wärme* 54, 270-82(1931)—An essential part of the heat loss of all furnaces is contained in the sensible heat content of the waste gases. This may be std. from the heat of combustion of the gases and a knowledge of the waste gases. The waste gas loss of the simple combustible gases, H₂, CH₄, ethylene, benzene, acetylene and CO, is computed for different conditions of temp., excess air and humidity. By means of these values the waste gas loss of com. gaseous fuels is obtained. Tables are given which show the heat content of the gases of combustion. An economical comparison is made of coal-, water-, and generator gas. ALLEN S. SMITH

Metallic ions as catalysts for the removal of sulfur dioxide from boiler-furnace gases. H. F. JOHNSTONE *Ind Eng Chem* 23, 559-61(1931)—The catalytic effect of small concns. of Fe⁺⁺⁺ and Mn ions on the oxidation of dissolved SO₂ materially increases the capacity of water for absorbing SO₂, and thereby decreases the vol. required for washing flue gases. At a concn. of 300 p. p. m. Mn, 3500 cu. ft. of an air mixt. contg. 0.325% SO₂ was washed with 1 gal. of water at an av. efficiency of 95.5% removal of SO₂. The concn. of H₂SO₄ attained was 29.6%. Fe⁺⁺⁺ requires concns. about 100 times those of Mn. Cu ions act as a strong inhibitor to the catalysis by Mn but not by Fe⁺⁺⁺. Zn, Ni and Cr ions have no effect either as inhibitors or as catalysts. When lime water is used for washing, CaSO₃ rather than Ca(HSO₃)₂ is formed. Four thousand cu. ft. of flue gas contg. 5% CO₂ and 0.1% SO₂ was washed per gal. of water contg. 250 p. p. m. Mn. A rotary scrubber was used and the time of contact was approx. 1/10 sec. This is equiv. to 1 1/4 tons of water per ton of coal. The efficiency dropped from 40% at the beginning to 10% at the end when the acid concn. was 3.7% H₂SO₄. Further work on the application of catalysts to the scrubbing of flue gases is in progress. H. F. JOHNSTONE

Stoichiometry of the blow. LLOYD LOGAN *Am Gas J* 134, 53-4(1931)—In the blow run of a water-gas set C is consumed by oxidation to CO or CO₂. There is a unique relation between CO, CO₂ and N₂ content of the blast gas, the compn. varying from 21% CO₂, 0% CO, 79% N₂ to 0% CO₂, 31.7% CO, 68.3% N₂. From the detn. of any one of the components the other 2 may be calcd., as well as the amt. of C consumed and the vol. of blast gas. Equations and charts are given for the calcn. of the unknown quantities from the analysis. CHANNING WILSON

Protection against fire hazards of gas appliances. F. C. VANDAYRE *J. Mines Gas* 55, 75-80(1931)—A review is given of results obtained in the Cleveland (O.) A. G. A. labs. B. J. C. VAN DER HORVEN

Deformation of bells and lifts of gas holders under the influence of gas pressure. M. PETTIPAS *J. Mines Gas* 55, 35-9(1931)—A review is given of the mechanics of water-gas holders. B. J. C. VAN DER HORVEN

The use of tar in the bell seals of gas holders. W. J. A. DUNSTON *Helv Gas* 51, 73-7(1931)—From a no. of drawings and from a pressure chart of the gas trapped under the seal of the telescope ring of a 2 lift holder in bottom position it is shown that even if the ring seal is originally filled with tar to the top a water layer will replace it partly as soon as it has been down to the bottom once. Only by lowering the water level in the foundation basin can a sufficient tar seal be maintained in the ring permanently. B. J. C. VAN DER HORVEN

The determination of sulfur in combustible liquids. I. F. GARELLI AND B. SALADINI *Ann chim opNicala* 21, 136-40(1931)—After reviewing the method of detg. S in illuminating oil by burning a weighed amt. and collecting the liberated SO₂ in alkali soln., a more general method is described suitable for crude as well as refined oils. This method is the common one using an O₂ bomb. A. W. CONTIERI

The plasticity of coal and the theory of coking. CLARENCE A. SKYLER *Proc S. Wales Inst Eng* 47, 9-14(1931)—When heated, coal deforms under const. pressure at a rate which diminishes as the contraction proceeds until it finally ceases, as a result, out of loss of the property of flow under pressure, but of the development of a resistance probably of the nature of a back pressure produced by the gases of decompn. The rate of deformation increases with pressure. With no pressure there is no flow. The phenomenon is one of plastic flow, not fusion. Plasticity involves 2 quantities, mobility (const. at a given temp.) and resistance (only conditionally const.). A plastometer

which permits calculation of the constants in c. g. s. units has been developed. In the coals investigated, when the rate of flow and temp are not too high, coal follows the laws of plastic flow. Softening point is the temp at which a min rate of flow is produced at a given pressure, such that a very small further contraction causes the resistance to equal the external pressure. The plasticity-temp relation is hyperbolic with 2 asymptotes, 1 approaching the temp axis and the other inclined to it. The plasticity of coal at the softening point is inversely proportional to the pressure. The plasticity does not increase indefinitely but reaches a max detd. by the thermal stability of the coal. This max decreases as the softening point of the coal increases and its thermal stability decreases. Preheating has a marked effect on plasticity. Coking is due to the plastic flow of the coal particles as a whole. The necessary pressure comes from the gases evolved.

Results of compressive strength, tumbler and shatter tests of foundry coke. H. A. J. PIETREAS *Helv. Chim. Acta* 51, 84-5 (1931).—The compressive strength was detd. on 10 cubical coke bodies sawed out and polished to 1 cc. For the tumbler tests 50 kg. coke was revolved for 4 min. in a drum, 1 m. long, 1 m. diam., 25 revolutions per min., then screened on 40 mm. round openings and the retained coke weighed. For the shatter test 25 kg. coke was dropped out of an Fe box with doubly hinged bottom halves on to a steel plate 1.80 m. lower. This test is repeated 4 times, the coke then screened on 80 mm. round holes and the retained coke weighed. From tabulated data of 52 expts. with 2 types of foundry coke it appears that the tumbler test gives the most consistent figures, the shatter test is more sensitive and less const., max. deviation for the former is about 3% of the average, for the latter around 10%. The compressive strength detn. has no practical significance for the valuation of the quality of the coke. B. J. C. v. d. H.

Heating inclined retorts with high B. t. u. gas. H. MÖLLER AND STIEF *Gas u. Wasserfach* 73, 1177-80 (1930).—Inclined retorts have been heated with mixts. of coke-oven gas with water gas, producer gas and flue gas, resp., as well as with coke-oven gas alone, which requires special burner provisions. The optional use of coke-oven gas increases the flexibility of gas manifold and permits a higher coke gas ratio when this is desirable. The authors state that all coal-carbonizing equipment should be designed to operate with high or low B. t. u. gas. R. W. RYAN

The geological history of coal (HICKLING) 8. Natural gas for metallurgical furnace (Gooss) 9. The origin of coal (FUCHS, HORV) 8. *Escoecaria agallocha* as a source of power alcohol (SEY, *et al.*) 22. Refractory materials (GEESEV) 19. Tower for effecting contact between gases and solid materials as in scrubbing fuel gas (U. S. pat. 1,798,307) 1. Working up oxidation products from paraffin, etc. (Ger. pat. 522,055) 22. Treating gas-washing oils (Ger. pat. 522,032) 22. Tower for liquid and gas contact operations such as washing fuel gases (U. S. pat. 1,798,563) 1.

ENGELDER, CARL J. Laboratory Manual of Gas, Oil and Fuel Analysis. New York: John Wiley & Sons, Inc.

FUEL. MADRUCK GAS FÜR SICHNELLE DRUCKENTWASSERUNG u. a. H. Ger. 516,761, Apr. 19, 1929. Crude peat is dehydrated by pressing with fully dried peat dust. The mass is heated during pressing and cooled by water-cooled plates.

Liquid fuels. SOCIÉTÉ NEUTRAL MOTOR FUEL, INC. Fr. 697,439, June 14, 1930. Motor fuels are improved by bringing them in intimate contact with a mixt. of an oxidizing agent sol. in water such as H_2O_2 and a compd. contg. Cl sol. in water such as $AlCl_3$ in the presence of water. A suitable app. is described.

Motor fuel. I. G. FARBERVIND A-G. Fr. 698,641, July 8, 1930. A valuable motor fuel is made by the thermal decompn. of mineral and tar oils in the state of vapor and preferably in the presence of catalysts at temps. above 600°. The decompn. takes place under conditions which cause the formation of large amts. of gaseous products, particularly unsatd. hydrocarbons, the appropriate fractions of which are sep'd. from the non-gaseous products.

Composition for treating fuel. FRED L. MENNIE. U. S. 1,799,631, April 7. A compn. for treating fuel such as coal-dust briquets in order to improve its burning properties comprises $KMnO_4$ 6, $KClO_4$ 12, "table salt" 32, crushed rock salt 50 parts, mixed with water.

Utilizing carbon saturated with oxygen. MAX PLATZCH. Ger. 506,274, Nov. 13, 1927. C satd. with O may be used, e. g., for the manuf. of CO relatively free from N as a motor fuel, for the direct internal heating of metallurgical, etc., furnaces, or for the

heating of closed pressure vessels used for cracking oils, etc. Few or no details of these applications are given

Destructive distillation of suspended fuel particles. FRIEDRICH BARTLING U. S. 1,798,903, March 31 Various details of app. suitable for use in the destructive distn. of powd. bituminous material are described

Rotary kiln suitable for low-temperature coal distillation, etc. EDMUND ROSE (to C. H. Illworth) U. S. 1,799,500, March 31

Low-temperature carbonization apparatus. FREDRICK L. DUFFIELD Fr. 698,609, July 4, 1930

Carbonizing retorts and furnaces. WOODALL DUCKHAM (1920), LTD., and ARTHUR McD DUCKHAM Fr. 699,551, June 30, 1930

Hydrogenation reactions. I. G. FARNENING A-G Fr. 697,539, June 17, 1930 In reactions using H₂ at a high temp. particularly in the hydrogenation of coal, tars, oils, etc., the reaction vessels are protected against the destructive action of the H₂ by covering their surfaces with an alloy of Cu and Zn Other metals such as Mn, Ni, Fe, Co, Al or Ag may be incorporated to increase the mech. resistance

Hydrogenating coal, etc. I. G. FARNENING A-G Fr. 698,425, July 5, 1930 The mixt. of oily and solid substances obtained by the hydrogenation under pressure of coals, oils, etc., is heated to a temp. higher than that used in the treatment until the solids are agglomerated into larger particles and sepd. Cf. C. A. 25, 2549

Hydrocarbons I. G. FARNENING A-G (Rudolf Wietzel and Bruno Engel, inventors) Ger. 522,463, June 10, 1927 In the manuf. of hydrocarbons, principally liquid, from coal, tars, mineral oils, etc., by destructive hydrogenation, cracking, or extrn. with org. solvents under pressure, a small quantity of an org. substance capable of lowering the surface tension of the products is present, e. g., an emulsifying agent such as Na isopropyl-naphthalene sulfonate Cf. C. A. 25, 1008

Hydrocarbons. I. G. FARNENING A-G Fr. 698,244, July 1, 1930 Valuable hydrocarbons are prepd. from coal, tars, mineral oils, etc., by treatment with H₂ or gases contg. It The finely divided starting material is allowed to fall freely through a vessel contg. heating elements and filled with H₂ under pressure

Recovery of phenols. F. RASCHIG & S. H. Fr. 698,300, July 4, 1930. Residual waters from coking plants, fuel-distn. app., etc., are obtained free from phenols by washing the distn. gases with solvents for the phenols at a temp. above the dew pt. of the mixt. of gases.

Acetylene cylinder filled with plastic material. BRUCE MILLER (to Air Reduction Co.), U. S. 1,799,803, April 7. Structural features.

Gas. THE POWER GAS CORP., LTD. and NIELS E. RAMMUSSEN. Fr. 697,788, June 23, 1930 An app. is described for making a mixt. of water gas and producer gas.

Fuel gas. JULIUS PINTSCH A-G Ger. 522,493, Mar. 27, 1929. The gaseous by-products (consisting mainly of CO) from the synthesis of NH₃ from water gas are mixed with water gas and waste gas, e. g., slue gas, in such amts. that a mixt. of the approx. compn. of generator gas is obtained

Fuel gas for automobiles. JEGOR I. BRONN and CONCORDIA-BERGHAU A-G Ger. 510,850, July 28, 1926. The gas consists of C₂H₄ mixed with gaseous hydrocarbons from the low-temp. distn. of coal or lignite, or from the cracking of mw. oils, etc.

Producer gas. LOUIS CHAVANNE. U. S. 1,799,885 April 7 In utilizing solid fuels in a slagging gas producer, the lower portion of the fuel bed is blasted with a gaseous mixt. of a free O content regulated for melting the fuel ashes in a fusion zone extending above the blasting level and maintained to a thickness not exceeding one-fourth of the total height of the fuel column, and for maintaining above this an extended zone of low-temp. distn. An aq. fluid is added near the top of the fusion zone, gases and entrained by-products are withdrawn, and molten products are removed from the base of the column U. S. 1,799,888 relates to a generally similar process in which fuel dust is introduced in the zone of fusion and fuel in lump form at the top of the fuel column App. is described.

Water gas. WM. J. EDMONDS (to Commercial Solvents Corp.). U. S. 1,799,359, April 7. In a cyclic process for the production of substantially S-free water gas, fuel is subjected to alternate air and steam blows, the water gas resulting from the steam blows is desulfurized by contact with solid desulfurizing material such as iron or Fe oxide and the desulfurizing material is revived at intervals, during air blows, by contact with gases comprising, at least in part, the gaseous products of the air blows. App. is described

Water gas. OESTERREICHISCH AMERIKANISCHE MAGNESIT A-G. Fr. 697,489,

June 16, 1930 Water gas poor, in or free from, CO is made by the action of steam on a mixt. of carbonaceous material and the oxide, hydroxide or carbonate of Mg at a temp. above the disson temp. of $MgCO_3$ but not appreciably passing 750° . CO may be eliminated from a gas contg. it by passing the gas over the above mixt. at a temp. not appreciably above 500° .

Oil gas. EDWARD A. DINTERLE U. S. 1,708,372, March 31 Hot combustion gases are supplied upwardly to and through vertically spaced banks of heating tubes stacked in a chamber, and oil is simultaneously sprayed on the outside surfaces of the heating tubes for generation of oil gas, and the oil gas thus formed is withdrawn from the bottom of the chamber. App. is described.

Desulfurizing gases. ISTVÁN HILYVÁDY and KARL KOLLER Fr. 698,327, July 3, 1930 Industrial gases are desulfurized with recovery of S, by bringing the gases to be purified into intimate contact with an aq. soln. of finely ground MnO_2 . The mixt. of S and sulfide formed is sepd. by filtration and the sulfide is transformed in known manner to hydrate and thence by the action of dry air to MnO_2 . The mass is dried and the S dissolved out. Cf. C. A. 25, 2550.

Removing hydrogen sulfide from gases. WINTERSHALL A.-G. Ger. 516,851, Sept. 17, 1927. The gases are passed into a suspension of basic ferric salt in a ferric salt soln., prepd. by oxidizing the corresponding ferrous salt with HNO_3 . Thus, a soln. of $FeCl_3$ in neutral or weakly acid soln. is oxidized by HNO_3 . Part of the $FeCl_3$ is pptd. as $FeCl_3 \cdot 6Fe(OH)_3$, the rest forming $FeCl_3$ which remains in soln. This liquid absorbs the H_2S in gases with disposition of S, which can be recovered by filtration and centrifuge.

Treating slurry from gas-purification products. FREDERICK W. SPURR, JR. (to Koppers Co.) U. S. 1,708,012, March 31. Slurry produced in the purification liquid is submitted to a flotation operation and S is removed as a froth concentrate, and metallic compds. are removed as tailings. An arrangement of app. is described. Cf. C. A. 24, 4922.

Oil-gas generator. EPOCH RECTOR (to Rector Gauthier Co.). U. S. 1,709,247, April 7 Various structural details are described.

Gas producers. CHARLES WHITFIELD Fr. 698,500, July 2, 1930 Means for regulating the combustion of the fuel is described.

Vertical gas-producer. JOHN U. McDONALD U. S. 1,708,617, March 31.

Distillation column for gas producers. CARL BLEYER, Fr. 698,042, June 13, 1930.

Means for recovering heat from flue gases. HARRY FAHRBACH Fr. 697,846, June 24, 1930.

Purifying tar acid-bearing oils. SOLOMON CAPLAN (to Combustion Utilities Corp.). U. S. 1,799,780, April 7 In removing color producing substances from tar acid bearing oils, the oil is distd., with refluxing and condensation, and, during one of the steps prior to condensation, is brought into contact with a non acid ferric compd. such as ferric stearate.

Coke ovens. AUGUST KOPFERS Fr. 697,625, June 18, 1930 The drawing mechanism and control are described.

Coke ovens. FRIEDRICH CARL STILL Fr. 697,618, June 18, 1930. The loading mechanism is described.

Regenerative coke ovens. HEINRICH KOPFERS A.-G. Fr. 697,472, June 16, 1930. Constructional details are given.

Horizontal regenerative coke oven, with vertical heating flues. KOKSOFENBAU UND GASVERWERTUNG A.-G. Ger. 504,113, Aug. 7, 1928.

Coke ovens with vertical chambers. N. V. SILICA EN OVENBOUW Mij. Fr. 697,680, June 19, 1930.

Coke-oven battery. THE KOPFERS CO. Ger. 522,388, Oct. 16, 1928.

Coking-retort oven. JOSEPH BECKER (to Koppers Co.) U. S. 1,709,670, April 7.

Coke-oven smoke-discharge system. JOSEPH BECKER (to Koppers Co.). U. S. 1,708,517, March 31. Structural features.

Fastening device for coke-oven door. JAMES H. GLEASON (to Gas Machinery Co.) U. S. 1,708,569, March 31.

Plant for cooling glowing coke. GEBRÜDER SÜLLER A.-G. Ger. 516,870, May 22, 1928. Details are given.

22—PETROLEUM, LUBRICANTS, ASPHALT AND WOOD PRODUCTS

W. F. PARAGHIE

Methods of utilizing for power purposes the various petroleum residuals and by-products. A. I. DUNSTAN *J. Inst. Petroleum Tech* 16, 701-22 (1930).—The practices of the petroleum industry in burning its acid sludge, coke, asphalt and refinery gases are reviewed. EMMA B. CRANDAL

Refrigerating requirements for oil refining. HARRIS FRUIT *Refrigerating Eng.* 21, 311-3 (1931).—The application of refrigeration to the removal of wax from paraffin-base lubricating oils is described. A. H. JOHNSON

Refrigeration in oil refining. J. M. WADSWORTH *Refrigerating Eng.* 21, 95-100 (1931).—The operation of an oil refinery using paraffin base crude oil is described. Wax is removed by gravity, by filtration or by centrifuging. A. H. JOHNSON

The action of ferric oxide and ferric chloride as catalysts in the cracking and hydrogenation of oil from Ragusa. L. SALMOIRAGHI *Ann. chim. applicata* 21, 27-37 (1931).—Samples of oil from Ragusa containing about 2.85% S and having a sp. gr. of 0.9667 were submitted to simple cracking by heating to 490° and also to cracking with hydrogenation in the presence of Fe_2O_3 and $FeCl_3$. The amts. of the chloride were varied from 5.97 to 11.07%. No gas is formed by simple cracking. With Fe_2O_3 , 17.3% was formed (mainly $C_{11}H_8$, $C_{12}H_8$ with 80% H_2), while with $FeCl_3$, 7.4 to 13.6% of gas was made. In the simple cracking, only 4.5% of the product distd. under 250°, and 41% of residue was obtained. With Fe_2O_3 , 37% distd. under 160°, 40% up to 250°, and the residue was 3.0%. With $FeCl_3$, 22% distd. below 150°, 53% below 250°, and there was no residue. With Fe_2O_3 , the S was about evenly divided between the distillates and the residue, whereas with increasing amts. of $FeCl_3$, up to 75.5% of it was retained in the residue. A. W. C.

Crude oil emulsions. A. TZATUR'YAN *Groznenskiy Neftyanik* 1, No. 2-3, 39-43 (1930-31).—The following methods for breaking up crude oil emulsions are described: (1) heating in open containers, (2) heating in closed containers, (3) elec. method (with a potential of 1000 v.), (4) the Tret O-Lite method, (5) centrifuging method, (6) combined methods, i. e., a combination of two of the methods mentioned above. A. A. BORISILINOK

Fractional distillation of crude oil. E. O. SLATER. *Oil Bull.* 17, 272-4 (1931).—The 4 point distn. method for evaluating crude oils is described in detail. It consists in a preliminary distn. (topping to a vapor temp. of 500°) under specified conditions of a weighed sample of the crude oil. Above 250° F. the distn. rate is not to exceed 10 ml. per min. A 300 ml. sample of the tops (contg. the gasoline and kerosene) from this distn. is rerun in the app. described in A. S. T. M. specification D-285-30-T, the condenser tube being packed in cracked ice and a dry graduate cylinder covered by blotting paper and immersed in water at 60° F. being the receiver. The distn. rate is 2-3 ml. per min. for the 1st 5 ml. distd. and 4-5 ml. thereafter. The vols. of distillate in ml. are detd. when the vapor temp. reaches the following points: 221°F., 284°F., 392°F., and 437°F. From these vols. the percentage of gasoline recovered at each of these temps. and the percentage of gasoline in the crude oil are calcd. as follows. The percentage of gasoline recovered at each distn. point is the vol. recovered divided by the vol. of tops distd. times 100, this is the percentage of tops. The percentage of tops is multiplied by the percentage of net tops calcd. from the crude distn. (from the net vol. of crude oil and the net tops or total tops minus water), and the resulting figure is multiplied by a different factor for each temp. 5, 2, 1.11 and 1, resp., for the 4 distn. points to give the percentage of gasoline in the crude oil. D. F. BROWN

The detonation in internal-combustion engines. K. SCHNAUFER. *Z. Ver. deut. Ing.* 75, 455-6 (1931).—S. uses a single cylinder engine in which the firing spark plug is located at one end of the cylinder head. Three other spark plugs are located along the direction of flame travel. These are connected to a source of high frequency current and to an oscillograph. The passage of the flame front ionizes the gases between the electrodes, thus allowing the passage of the current and giving a record on the oscillograph. The firing is also recorded, as well as the pressure-time curve, and the oscillations of a timing fork. In the absence of detonation, the flame travels at an increasing speed, covering the 13 cm. of travel in 0.005 to 0.01 sec. When the engine is made to detonate, the flame reaches the last 2 indicators practically simultaneously. This shows the detonation consists in the spontaneous ignition of the whole mass of unburned gas compressed in front of the flame. G. CALINGBERT

Detonation characteristics of some aliphatic olefin hydrocarbons. WHEELER, G.

LOVELL, JOHN M. CAMPBELL AND T. A. BOYD *Ind Eng Chem* 23, 855-8(1931) — The relative knock ratings of 25 aliphatic olefin hydrocarbons in admixt with gasoline have been detd. With the C_4H_{10} equiv. as a basis of comparison, the straight-chain olefins showed that (1) in homologous series there is increased knock tendency with increased length of chain; (2) there is a progressive decrease in knock tendency in isomeric series, with centralization of the double bond and (3) tendency to knock is roughly detd. by length of the longest satd. C chain. In branched-chain olefins the tendency to knock decreases upon the introduction of a double bond, both the position of the bond and the branched structure of the mol. are important. Cf. C. A. 25, 1367. J. L. FISKE.

The behavior of antiknocks. A. EGERTON AND L. M. PIERCEON *Nature* 127, 591(1931) — $PbEt_4$ is introduced into an evacuated vessel which is then heated to 235° . A charge of C_4H_{10} and air is then introduced and exploded. The $PbEt_4$ does not slow down the explosion and may even accelerate it. If the expt. is repeated in the presence of a small quantity of air, the well-known retarding effect of Pb is observed, showing that Pb must be oxidized before it becomes an effective inhibitor. G. CALINGAERT.

A study of the knock in internal-combustion engines. L. AUEA *Forsch Geb Ingenieurw.* Ausgabe B 2, *Forschungsheft* 340, 18 pp (1931) — After a review of the literature, A. describes his own expts. The engine used is a single-cylinder, 4-cycle, water-cooled, valve-in-head Baumann-Muenchen engine, 6 h p, 165 mm bore, 180 mm stroke and 320 r p m. The compression ratio is varied from 3.6 to 10.1 by varying the length of the connecting rod. Two types of knock indicator are used. The 1st is derived from that described by Polster (*Forsch Arb Geb Ingenieurw.* 172(1915)). This instrument, which is bolted to the underside of the piston, functions through the acceleration imparted to the piston by the knock. It comprises an "accelerometer," which is held in position by a spring, and an adjustable electromagnet. By varying the exciting current of the magnet, it is possible to balance the force so that the mass is just deflected from its normal position by each detonation. This slight deflection is detected by the breaking of an elec. contact. The actual intensity of the knock, expressed in kg/sq. cm, is calcd. from the mass of the piston and the accelerometer and the force of the spring and the magnet. The time of its occurrence is detd. by recording the time of breaking of the contact. The 2nd knock indicator has a small piston, the head of which is brought flush with the wall of the combustion chamber. The piston is held in position by a strong spring. The tension on the spring is adjusted to balance the pressure of the knock, thereby preventing a deflection of the piston. The pressure is calcd. from the known balancing pressure which holds the piston in position. The fuel used is "Dapohn" gasoline of d. 0.711. The curves of "intensity of knock" and of "time of occurrence of knock" versus air-fuel ratio pass through a max. and a min., resp., around the point corresponding to the theoretical air-fuel ratio. With a compression ratio of 6.1, 10° spark advance and 310 r p m., the following typical data give successively the air-fuel ratio (by wt.), the intensity of knock (kg/sq. cm), the time of knock (degrees of revolution after the top dead center): 10.8, 2.66, 33.5, 14.4, 7.40, 19.9, 15.7, 2.63, 40.0. At 6.1 compression and 310 r p m., advancing the spark from 0 to 20° increases the intensity of the knock from 0.75 to 6.74 kg/sq. cm, while the time of the knock advances from 69.8 to 17.3° after the top dead center. At 6.1 compression and 10° spark advance, varying the speed from 205 to 336 r p m. decreases the knock from 8.61 to 2.88, while the time of knock is retarded from 21.6 to 34.6° after the top dead center. Recalcd. in sec. instead of r p m., the latter data show a practically const. lapse of 0.022 sec. between ignition and detonation, irrespective of the speed. The greatest effect on detonation is shown by the variation of the compression ratio. At 296 r p m. and 15° advance, ratios of 5, 6 and 7 to 1 give intensities of 5.12, 7.19 and 24.08, resp., with corresponding times of 34.3, 21.2 and 14.2° after the top dead center. At 6.1 compression, 10° advance, 310 r p m., air intake temps. of 23, 49 and 74° give intensities of knock of 1.69, 2.66 and 6.50, resp., with corresponding timings of 45.9, 26.3 and 24.6° after the top dead center. Finally, the knock was found to be unaffected in intensity and timing by variations in intake manifold pressure from 44 to 68 cm. of water. A comparison of these results with published data on the rate of flame propagation in explosive mixts. led A. to believe that detonation is caused either by an explosion wave or by the spontaneous explosion of the unburned gases compressed adiabatically by the flame front. G. CALINGAERT.

Improved bearing lubrication. HENRY BAILEY *Bull. soc. encour. ind. nat.* 130, 95-108(1931) — Lubrication of bearings by oil under pressure results in decreased power consumption and metal wear.

Extracts obtained in the manufacture of ethereal oils from conifers. P. A. BORISOV. *Trudy Vyskuzh. Nauch. Issledovatel'skii Inst. Khim. Promy.* 4, 78-86(1923) — The mixt.

L. W. T. CUMMINGS

under investigation is a by-product in the manuf. of fir-tree oil obtained by treatment of conifers with steam. It is sol. in water and has marked tanning qualities. After hydrolysis continued for 80 hrs. ether extd. from this soln. phenols, catechol and a solid of sweet taste. The rest of the aq. soln. gave reactions of aldehydes, reacted with phenylhydrazine, reduced Ag and Cu salts, was optically active and was sweet in taste. It is regarded as a carbohydrate. From results of detns. of carbohydrates with Fehling soln. it was concluded that its liberation from a higher complex is due to hydrolysis. Another product of hydrolysis was a ppt. from which ether extd. phenols (catechol and resorcinol), acids (protocatechuic and gallic) and a neutral substance. The tanning qualities of the exts. were tested by reactions with lead acetate, gelatin and glue and powder lude. As the results showed tanning qualities of the catechol group, actual tanning of pieces of hide was undertaken by shaking them with 0.2-4% solns. of the tanning substances, treatments with hot water and 5% AcOH showed successful tanning.

J. G. TOLPIN

Recovery of organic nitrogen by dry distillation of agricultural waste. V. DIRCO. *Affidavit* 10, 97-101(1930).—Niff, was obtained in 2.7% yield, together with tar, charcoal, AcOH and MeOH.

B. C. A.

Excoecaria agallocha as a source of power alcohol. HEMENDRA KUMAR SEN, SINHU BRUSAN GHOSH AND PATIT PABAN PAL. *Proc. 15th Indian Sci. Cong.* 1928, 183.—The sawdust from *Excoecaria agallocha*, which grows abundantly in Sundarban tracts and is now being used plentifully for the match industry and for making cheap packing cases, was investigated. While with most species previously examd. 22-24% of reducing sugar was the max. production after hydrolysis by Simonsen's process, namely, digestion with very dil. H_2SO_4 under 0 atm. pressure, the yield of sugar was about 40% of the wt. of dry sawdust with *Excoecaria agallocha*. Of this about 70% was readily fermentable, giving approx. 39 gal. of abs. alc., a figure practically double that which was obtained in the past. With appropriate figures of cost, a gal. of abs. alcohol from *Excoecaria agallocha* would work out at 6.03 annas. Given a continued supply of this wood, there is thus an excellent prospect of a power alc. industry in the province of Bengal.

E. J. C.

The utilization of waste wood (SCHMIDT) 23. "Liquid rosin" and its possible uses (SCHMIDT) 23. Rayon oils (MULLIN, CALDWELL) 23. New compounds of Ti (GARDNER, BIELOUS) 10. The use of Al for oil-lease tanks (SCHMIDT, *et al.*) 9. Hydrocarbons (Ger. pat. 522,463 and Fr. pat. 698,214) 21. Motor fuel (Fr. pat. 698,611) 21. Fuel gas for automobiles (Ger. pat. 516,850) 21. Hydrogenating coal, etc. (Fr. pat. 698,425) 21. Oxidizing paraffins, etc. (Ger. pat. 622,361) 10. Oil filters (Fr. pat. 697,919) 1. Hydrogenation reactions (Fr. pat. 698,539) 21. Uninflammable solvent containing hydrocarbons (Fr. pat. 697,600) 13. Tetraalkyl 1'b (U. S. pat. 1,708,693) 10.

Groznenakli Neftyanik (Grozny Petroleum Worker) (*New Journal*). Published monthly by Chechensku Oblastnoi Soyuz Gornorabochikh i Trast Grozneft, Prospekt Revolyutsii 24, Komnata 33, Grozny, U. S. S. R. Vol. 1, No. 1 appeared in Nov., 1930. Price 3 rubles.

TAPP, HARRY F. Handbook of Oil Burning. New York: Am. Oil Burner Assoc. 629 pp.

Tank and breather to prevent loss of vapors from volatile liquids such as petroleum fractions during storage. HAROLD V. ATWELL (to Standard Oil Co. of Ind.) U. S. 1,800,013, April 7.

"Tube bundle" heat-exchange apparatus for use in refining petroleum oils. CHARLES H. LEACH. U. S. 1,799,471, April 7. Structural features.

Treating petroleum residues. PERCH M. TRAVIS (to Travis Process Corp.) U. S. 1,798,311, March 31. For removing solid and semi-solid materials from petroleum residues, the residual oil is heated to about 50-55° and centrifuged at this temp. and the thus-sepd. solid material is conveyed axially in one direction in the sepg. zone while the liquid moves in the opposite direction, and the materials are continuously discharged. Various details of app. and operation are described. Cf. C. A. 24, 234.

Composition for increasing the conductivity of hydrocarbons. WOLF KRITCHEVSKY and ELLIOTT MORRILL. Can. 309,980, Mar. 31, 1931. A compn. for increasing the cond. of light petroleum distillate consists of a partially or completely substituted ammonium salt of an aliphatic acid, e. g., triethanolamine oleate and a solvent. The solvent is not necessarily a petroleum hydrocarbon, it may be propyl alc., ethyl acetate, acetone, etc.

Catalytic cracking of petroleum oils. FRANK B. MILLER (to Silica Gel Corp.). U. S. 1,799,858, April 7. Oil to be cracked is vaporized and there is fed into a stream of the vapor a pulverized catalyst consisting of silica gel and an active material such as one containing Fe and the vapor with the catalyst in suspension is passed through a zone of suitable temp and the catalyst is then sep'd. for further use. App is described. A similar procedure may be used in catalyzing various other reactions.

Cracking hydrocarbon oil. CARLOS P. DUBBS (to Universal Oil Products Co.). U. S. 1,799,218, April 7. Fresh untreated oil is introduced to one end of a series of externally unheated zones, and is caused to pass successively through this series of unheated zones. Progressively higher pressure is maintained on the succeeding zones of the series, and unvaporized oil is taken off from the last zone of the series, and the oil is subjected to cracking conditions of heat and pressure and the heat for this purpose is supplied exteriorly of the zones mentioned. After removal of heavier constituents, the heated conversion products are introduced into the zone of the series from which the unvaporized oil is taken off and vapors released from the oil are caused to move through the series of zones countercurrent to the movement of the oil, and a phys. commingling of the vapors and oil in each of the zones is effected and vapors which have escaped condensation are removed from the zones of the series into which the fresh oil is introduced and are subjected to a final condensation, and the resulting distillate is collected as the ultimate product of the process. App is described. Cf C. A. 25, 587.

Cracking hydrocarbon oils. GASOLINE PRODUCTS CO., INC. Fr 697,397, May 30, 1930. In cracking hydrocarbons a cold hydrocarbon is injected into the final and hot test portion of the cracking worm to prevent supercracking and an excessive deposition of C, while the injected oil is cracked to a certain degree. The injected oil may be of the same characteristic as the oil to be cracked. A suitable app is described.

Cracking oils. ROBERT T. POLLOCK (to Universal Oil Products Co.). U. S. 1,799,530, April 7. Oil is passed through a heating zone such as a pipe coil and thence to an expansion chamber. Vapors are taken from the expansion chamber to a reflux dephlegmator in which they are cooled by a regulatable supply of cool liquid such as water and the remaining vapors are led to a condenser. Raw oil is supplied to the condenser to condense the vapors and preheat the raw oil and the preheated raw oil is mixed with the reflux condensate from the dephlegmator and the mixt., free from unvaporized residue, is led under pump pressure to the heating zone (the latter being maintained under substantial superatm. pressure and the expansion zone being under less pressure, as are also the dephlegmator and condenser). Cf C. A. 25, 2279.

Cracking oils. PETROLEUM CONVERSION CORP. Fr 697,817, June 23, 1930. Petroleum oils are transformed into oils suitable for motor fuel by heating a gaseous vehicle to a temp at least equal to the temp of transformation, passing it into a reaction zone, submitting the petroleum oil to such conditions that it vaporizes without any appreciable cracking and mixing the gas and vapor in the reaction zone to provoke cracking of the vapor. An app is described. Cf C. A. 25, 2844.

Oil-cracking apparatus. LYMAN C. HERR (to Universal Oil Products Co.). U. S. 1,799,234, April 7. Metal walls of oil-cracking app are provided with a metal lining which is made in a plurality of sections, each anchored to the wall at a central point of the section, so that the sections expand from their centrally anchored points. Expansion folds formed from a different metal are interposed between and at all sides of the sections.

Plant for cracking hydrocarbon oils. EDWARD E. STEWART U. S. 1,798,338, March 31. Numerous structural features are described.

Hydrocarbon oil purification and refining. JACOB C. MORRELL (to Universal Oil Products Co.). U. S. 1,799,431, April 7. Oils such as S-contg. "cracked distillates" are heated with added metal such as Zn and with glacial HOAc, simultaneously, to convert refractory S compounds in the oil.

Tube-bundle heat-exchange apparatus suitable for heating hydrocarbon oils. KENNETH B. RIS (to Griscom Russell Co.). U. S. 1,798,354, March 31. Structural features.

Treating hydrocarbon oils to effect distillation. JACOB B. HERR (to Universal Oil Products Co.). U. S. 1,799,231, April 7. Oil is fed into the portion of an enlarged chamber where vapor sep'n takes place, and the unvaporized portion of the oil is caused to have a back and forth cascading descent in the chamber by which the oil is broken up and vapors are evolved. Unvaporized portions of the oil are collected as a body in the lower portion of the chamber, and the ascent of released vapors is accelerated by causing them to ascend through passages of successively increased cross-sectional area. App is described.

Use of mercury in distilling hydrocarbon oils. ARTHUR E. PEW, JR. (to Sun Oil Co.) U S 1,709,640, April 7. A stream of Hg vapor and a stream of relatively cold liquid oil are flowed into a confined path where they intimately mix and the mixt is passed into a relatively less confined space, the temp of the Hg vapor being sufficiently higher than that of the oil that, by heat exchange and condensation of Hg vapor, a portion of the oil will be vaporized, and the condensed Hg is sep'd from the unvaporized oil. App is described.

Distillation of hydrocarbon oil. JOHN S. WALLIS (to Foster Wheeler Corp.) Can 310,658, Apr 21, 1931. Hydrocarbon oil having a high b p is blended with oil having a lower b p and the mixt heated to a temp to vaporize the lower boiling constituent, the partial pressure effect of the vapors formed is utilized to cause vaporization of the higher boiling constituent, and the vapors are introduced into a fractionating tower and the fractions sep'd.

Distilling heavy hydrocarbons. SOC DES ÉTABLISSEMENTS BARNET Fr. 698,744, Oct 14, 1929. An app is described for distg or cracking very heavy hydrocarbons by bubbling them through molten Pb.

Device for measuring and sampling fluids such as oils in oil fields. HOLMES H. DYER. U S 1,798,788, March 31. Structural features.

Oil-shale distillation retort. CLAYTON O. WHITE. U S 1,799,268, April 7. Various structural details are described of a retort formed with sections having inclined bottoms over which powd oil shale will pass by gravity.

Converting oils into products of lower boiling point. CARBON P. DUBBS (to Universal Oil Products Co.) U S 1,799,413, April 7. An app is described comprising a heating zone such as a pipe system through which the oil is passed and which is connected with inclined vaporizing tubes which in turn are connected to collecting tanks adapted to permit settling of the cracked oil, and connections are provided for drawing off the residual product and for returning the lighter ends to the heating zone for re-treatment, and for controlling the relative temps of the heating and vaporizing zones.

Fractionating mineral oils. ALFRED R. EARL and THOMAS W. REEVES. U S 1,799,414, April 7. A series of interconnected units such as stills is maintained under progressively varying pressures (some above atm pressure), and an intermediate unit of the series is at atm pressures and some other units are below atm pressure. Oils are introduced into the system by means of the sub-atm pressures in one portion of the series of units and superheated steam under pressure is permitted to flow through the first mentioned series of units and to expand in the units under sub-atm pressure, and the pressures of the steam in the different units are utilized to maintain levels of oil at variant altitudes in the unit system, certain fractions are removed from the unit of greatest vacuum. Numerous details of app and operation are described.

Colored mineral oils. PATENT FUELS & COLOR CORP. Fr 698,153, June 27, 1930. Mineral oil (gasoline) or other petroleum distillates are colored in a stable manner by dyes and a stabilizing agent of a colloidal nature which may be a metal compd of a higher fatty acid, e g, Al palmitate, Mg oleate, Zn stearate. The dyes may be arylated rosanilines, basic triphenylmethanes, indulines, indazines, rhodamines, etc. If the dye has a strongly basic nature, an org acid such as stearic acid or benzoic is added if necessary.

Recovery of volatile substances such as gasoline from gases. HENRY J. NICHOLS, JR. and ERIC W. LUSTER (to Standard Oil Development Co.) U S 1,799,619, April 7. A low-pressure still gas or the like is compressed by the expansion of a high pressure gas such as a pressure-still gas and the compressed gas is cooled. The gases are mixed at an intermediate pressure, and volatile substances are recovered from the mixed gases. App is described.

"Sweetening" gasoline. HUGH H. CANNON (to Cannon-Prutzman Treating Processes, Ltd.) U S 1,798,784, March 31. Sulfurous petroleum products are "sweetened" by use of a dry powd reagent comprising PbO, an alk. earth metal hydroxide, an alkali metal hydroxide and a porous earth such as diatomaceous earth. Cf C A 25, 1067.

Working up oxidation products from paraffin, etc. I. G. FARBER and A. G. (Christoph Beck and Franz Kremp, inventors) Ger 522,055, May 27, 1928. To isolate the acids from the oxidation products of solid hydrocarbons such as paraffin and montan wax, the products are treated with NH_3 at atm or raised pressure and with or without simultaneous or subsequent addn of water. The acids are thus converted, according to the conditions, into amides or NH_4 salts. The amides may be freed from unchanged products by extg the latter with petroleum ether. The NH_4 salts sep out during the

reaction. The amides or NH_4 salts are converted into acids in known manner. Examples are given. *CI C A* 24, 4792.

Treating gas-washing oils. FRITZ SEIDENSCHNER. *Ger* 522,032, June 30, 1923. To sep. benzine from heavy oil that has been used to wash the distn. gases of lignite, the heavy oil contg. benzine is run on to the surface of a body of benzine free heavy oil heated to 150-200° in a gas heated vessel fitted with a dephlegmator. Benzine-free heavy oil is continuously withdrawn from the vessel.

Oil filter. CHARLES W. MCKINLEY (to A C Spark Plug Co.). U. S. 1,798,947, March 31.

"Self-washing" oil filter. JOE W. MORRISON (to General Motors Corp.). U. S. 1,798,950, March 31.

Lubricating oils. MARTIAL BEQUIER. *Fr* 698,737, Oct. 14, 1929. Lubricating oils are regenerated during use by adding a small quantity of a homogeneous mixt. of a fixed oil, a fatty acid and a mineral oil.

Lubricating oils. F. HAUSE. *Hung* 102,489, May 14, 1930. Addn. to *Hung* 99,915. Previously neutralized S or S-contg. material is added to vegetable or mineral oils by chem. or mechanical methods.

Hydrocarbons of high boiling point range suitable for use as lubricating oils. RUDOLF WITZEL, WALTER SPEFFER and HANS KAEBLER (to I G Farbenind. A-G). U. S. 1,798,288, March 31. A mixt. of oxides of C and H is subjected to a treatment with a catalyst contg. in addn. to a metal of group 8 of the periodic system such as Fe and Co, a quantity of an alkali metal compd. such as Na having an alkyl not substantially greater than that of Na_2CO_3 and comprising less than 0.5 part (calcd. as alkali metal) per 100 parts of the group 8 metal; the treatment is effected at temps. of 270-350° and the resulting hydrocarbons of low b. p. range are subjected to a treatment under pressure at temps. below 250° in the presence of an inorg. anhyd. halide having a condensing action such as AlCl_3 which when treated with water causes strong evolution of heat.

Purifying mineral insulating and lubricating oil. FRANK M. CLARK and ARTHUR T. HARDING (to General Elec. Co.). U. S. 1,794,948, March 31. Oil contg. dissolved impurities such as that used in elec. transformers is extd. with triethyl phosphate to remove the impurities. *CI C A* 24, 5477.

Lubricating greases. JOSE DOBOS. *Fr* 697,770, June 21, 1930. Olive oil residues from olives, etc., are neutralized by means of a virgin rosin and a carbonate or lime.

Hydrogenating oils. STANDARD OIL DEVELOPMENT CO. *Fr* 697,524, June 16, 1930. Lubricating oils are prepd. by submitting a heavy hydrocarbon oil to a treatment with H under a pressure above 25 kg. and at a temp. above 371°, the formation of oil boiling below 204° being limited to below 20% of the amt. of oil charged, so that the impurities of the oil are eliminated or rendered inoffensive. An app. is described. *CI C A* 25, 808.

Hydrogenating oils. STANDARD OIL DEVELOPMENT CO. *Fr* 697,724, June 20, 1930. Lubricating and other oils are prepd. by submitting heavy hydrocarbon oil to the action of a gas rich in free H, at a high temp. (371-454°) and under high pressure. A part of the oil in the liquid state is withdrawn and the pressure of the withdrawn part reduced so that a large proportion is vaporized. The vapor is separately condensed. An app. is described.

Condensation products. I G FARBEIND. A-G. *Fr* 697,700, June 20, 1930. Condensation products, which are either lubricating oils or resins, are made by condensing satd. or unsatd. fatty oils or waxes contg. OH groups with colophony or other resin acids.

Bitumen. SIEGFRIED HAMBURGER. *Fr* 698,081, June 26, 1930. Bitumen is produced from the acid and resinous by-products obtained in the refining of crude oil and oil residues by heating them to 80-100° and allowing them to stand until 3 layers are formed, the lower being H_2SO_4 , the next bitumen and the top oil.

Emulsion of tar or asphalt. HUGO NOVÁK. *Fr* 698,554, June 20, 1930. In emulsifying tar or asphalt, org. compds. contg. one or more atoms of Cl in the mol. such as CCl_4 , $\text{C}_2\text{H}_5\text{Cl}$, chloronaphthalene or Cl derivs. of oils are used as emulsifying agents in addn. to the usual stabilizing and emulsifying agents such as glue, casein, resin and naphthenic acids.

Deodorizing and decolorizing wood-distillation oils. DEUTSCHE GOLD- UND SILBER-SCHNEIDANSTALT VORM ROESSLER (Wilhelm Querfurth inventor). *Ger* 522,254, Mar. 4, 1927. The oils are vaporized and passed with H, at a temp. above

conclusion drawn by Bréguet from the study of fractional pptn. of such solns (Thesis, Lyon, 1924, p. 93). Other investigators who studied the fractionation of cellulose esters by ultrafiltration do not mention the H_2SO_4 contents of the products with which they worked. Specifications of max. limits for total sulfate and ash are entirely inadequate to insure proper quality of cellulose acetate; it is essential to specify a min. decompn. point (C. A. 19, 171), which should preferably be not below 195° . When the decompn. point is specified, total and neutralized sulfate are of secondary importance, as the decompn. point depends directly on these 2 values. The max. total H_2SO_4 content of 0.6% in cellulose acetate specified by the Fr. aeronautical service is practically meaningless, as cellulose acetate with this H_2SO_4 content can be highly unstable, and moreover cellulose acetates are being produced commercially which contain as little as 0.05% H_2SO_4 and 0.2% ash. The normal water content of cellulose acetate at 15° at a relative humidity of 50% is about 6%. In the detn. of Ac by sapon. with excess alkali and back titration with acid, the error introduced by the presence of H_2SO_4 is negligible in finished acetates, in the investigation of the mfg. and aging processes, in which total combined H_2SO_4 frequently reaches 2-3% and even more, the error is quite large and would lead to the conclusion that the Ac content decreases, contrary to C's results (C. A. 19, 171). Previous results have shown that nitrocellulose and cellulose acetate are easily washed free from acid, and the apparent difficulty reported by various authors is due to slow hydrolysis of the sulfo-cellulose esters in the presence of water, with liberation of free H_2SO_4 . A sample of com. nitrocellulose taken immediately after titration was divided into 2 portions, one was stabilized by washing 8 days in hard water (Rouen city water) (C. A. 18, 3715) and contained 1.13% H_2SO_4 , the other was stabilized by prolonged boiling in H_2O contg. a trace of HCl and contained 0.17% H_2SO_4 . They were subjected to the same fractional pptn. treatment, 400 cc. of a 5% Me_2CO soln. was used with the following results:

	Nitro contg. 1.13% H_2SO_4		Nitro contg. 0.17% H_2SO_4	
	Wt. of ppt.	H_2SO_4 , %	Wt. of ppt.	H_2SO_4 , %
Addn. of 140 cc. of 50% Me_2CO	4.8 g.	0.503	14.0 g.	0.503
Addn. of 30 cc. distd. H_2O	7.4		5.4	
Addn. of 20 cc. distd. H_2O	2.7	1.114	0	
Addn. of 50 cc. distd. H_2O	0		0	
Residue on evapn.	5.1	1.536	0.6	

These results show conclusively that, if the combined H_2SO_4 is not the controlling factor in the fractional pptn. of nitrocellulose, it plays a very important part. Though these results are the reverse of those obtained by Bréguet (Thesis) they do not necessarily contradict them, as B. used CaH_2 to ppt. Me_2CO solns. of celluloid. A. P. C.

Cellulose acetate and its production in Russia. CARL PETERS. *Kunststoffe* 21, 85-9 (1931).—Although the production of cellulose acetate elsewhere amounts to thousands of tons annually, the Russian industry is still in the early exptl. stages.

J. W. FERRY

Effect of neutral salts on the rate of hydrolysis of cellulose acetate in acetic acid solutions. J. T. PURSS AND C. J. STAUD. *J. Am. Chem. Soc.* 53, 1934-41 (1931).—The Na and K nitrates, sulfates and chlorides appear to exert a retarding influence on the rate of sapon. of cellulose acetate dispersed in $AcOH$. A greater effect was observed when nitrates and sulfates were used than when chlorides were employed. Addn. of neutral salts after the reaction had proceeded for 42 hrs. showed an effect similar to that noted when they were present at the beginning of the sapon.

C. J. WEST

Cellulose furoate, KENNETH A. KORA AND RALPH E. MONTONNA. *J. Am. Chem. Soc.* 53, 1889-91 (1931).—Cellulose di- and tri-furoates have been prepd. by treating standard cellulose with furoyl chloride in the presence of C_2H_5N . The esters were fibrous, dark colored and insol. in the usual solvents for cellulose esters. A considerable degradation of cellulose takes place during the reaction with the production of furoylated degradation products of cellulose which are black amorphous substances. The analysis of cellulose trifuroate by sapon. shows very clearly the fundamental difference of 2 of the alc. HO groups in the hexose residue from the third or primary HO and thus confirms the present accepted structure of the hexose residue.

C. J. WEST

Note on the cellulose amyloid. JUSTIN MUELLER. *Bull. soc. chim.* 47, 1400-4 (1930).—When cellulose is treated with H_2SO_4 (sp. gr. 1.58) and the resulting soln. dild., an amyloid seps in finely divided form. This amyloid gives a blue color with I_2 sol. in alkali, coagulates on drying, and does not reduce Fehling's soln. When dry it is less reactive than when wet. On these grounds it must be differentiated from hydro-cellulose and starch.

R. H. DOUGHY

The highest degree of nitration of cellulose. O. DU HURON, *Z. ges. Schiem-Sprengstoff* 26, 81(1911), cf. *C. A.* 25, 1671. By means of a process which he is not yet willing to describe, he has prepared nitrocellulose of any desired degree of nitration from 2 to 14% N. He believes the highest nitration product to be a homogeneous product. It is not decomposed by boiling water nor by very dil. acid or alk. solns. Its only impurity in 1911 is less than 2%.

The structure of nitrated cellulose. I. The swelling and disintegration of ramie cellulose in nitrating acids. FRANK D. MITCHELL AND MARJORIE MITCHELL, *J. Phys. Chem.* 34, 2589-2596(1930). Ramie fibers were nitrated to N contents ranging from 2 to 12% with various mixts. of H_2SO_4 and HNO_3 , the latter at various dilns. Phys. changes occurred at a N content of 7.5% which were probably assoc. with some structural change of the nitrated cellulose. The phys. changes appear to depend on N content rather than the acid compn. II. The x-ray examination of nitrocellulose. FRANK D. MITCHELL AND JAMES CRAIG, *Ibid.* 2507-21. The x-ray diagrams of nitrocellulose contg. up to 7.5% N were identical with that of cellulose, which is possibly due to intermolecular nitration. A N content of 7.5 to 10.4% resulted in phys. disintegration of the fibers and indefinite diffraction but the parallel arrangement of the chains in the nitrated structure was thought to persist since the regenerated cellulose was in an ordered form. At 10.3% N the stage of internal disorganization is passed. At 12.7 or 12.8% N there is a discontinuity and after this the comparatively definite trimutate is reached. I. A. S.

The structure of nitrocelluloses. DISMAREUX AND MATHIEU, *Compt. rend.* 192, 231-6(1931). X-ray diagrams of cotton cellulose nitrates contg. 11.33-14.0% N all showed distinct circles of diams. indicating regular distances of 0.6 to 7.5 Å. II., the increase being regular. The trimutate showed 2 intense circles corresponding to regular distances of 4.5 and 4.0 Å. resp. The product with less than 11.0% N showed 2 larger circles of lesser intensity, the one contg. 12.10% N having 2 concentric uniform regions which met. An examn. of the products by means of the polarizing microscope showed that the coloration of the fiber surfaces was that of a trimutate whereas the interior was characteristic of a lower nitrate, indicating heterogeneity. *Ibid.* 251-5.—When Miller diagrams of long chain paraffins, indicating that the different reticular planes admit of considerable variation in arrangement as respect to variation in temp., were compared with those of D and M, evidence of an anisotropic structure of nitrocellulose was observed. The interior circle of the Debye-Scherrer diagram of nitrocellulose of 14.0% N corresponds to the intense equatorial spots near the center of nitrocellulose, such spots caused by reticular planes not coinciding with the periodicity in the fiber. It is such a family of planes whose distribution remains periodic when the % of N is varied and which suddenly disappears when the nitrocellulose passes to the amorphous state.

Now layer-rolling presses for working cellulose. A. HANKE, *Kunststoffe* 21, 73-5 (1931), cf. *C. A.* 24, 6000.

New investments on viscose. HANNS SCHMIDT, *Kunststoffe* 21, 81(1931), cf. *C. A.* 24, 3012.

Production of viscose rayon with high tensile strength. HANNS SCHMIDT, *Chem. Ztg.* 55, 265-7, 280-8(1931)—A review.

The production of viscose artificial silk. D. I. PELLATT, *Textile Manufacturer* 55, 275-6, 312-3, 314-7, 307-9, 437-5(1930), 56, 31-3, 63-3, 103, 110, 137-8, 177, 188, 211-7, 257-9(1930)—Many details of the methods of mfg. viscose rayon are discussed.

Same factors in the prospects for staple fiber rayon. D. I. PELLATT, *Textile Manufacturer* 55, 475(1930)—An address.

The hollow rayons and synthetic yarns. CHAS. E. MULLIN AND FLORENCE H. CADWELL, *Textile Colorist* 52, 317-9(1930), cf. *C. A.* 23, 6310, 24, 4150—A review of the patents covering the manu. of the hollow yarns.

Rayon oils. II. CHAS. E. MULLIN AND FLORENCE H. CADWELL, *Textile Colorist* 52, 375-8, 414(1930), cf. *C. A.* 23, 1500—A description of the oils for, and their application to, the synthetic yarns, with a review of the literature and patents.

Effect of mild heat treatments on the chemical composition of wood. I. F. HANLEY AND JAN WIKSTRAK, *Ind. Eng. Chem.* 23, 181-3(1931). The chem. characteristics of samples of white ash and Sitka spruce wood were detd. before and after heating in closed tubes at 138° for 2, 4 and 8 days. Losses of pentosans and AcOH occurred in the ash wood whereas in the Sitka spruce the losses were hexosans and stable cellulose. The methoxyl content of both woods was practically unchanged. Indications of a change of carbohydrates to a lignin-like substance were observed.

The utilization of waste wood. W. SCHMIDT, *Papier-Fabr.* 20, Tech. Wiss., Tell

84-5(1931) —The combination of chem. softening and mech. disintegration, for making coarse pulps, is considered. R. H. DORTCH

Purified wood fiber. I. Physical and chemical properties. Geo. A. Richter. *Ind. Eng. Chem.* 23, 131-9(1931) —Various tests used to characterize paper pulps are described and interpreted. Certain chem. constants of 51 samples representing 11 kinds of pulp are reported. II. Paper-making characteristics of wood fiber high in α -cellulose. *Ibid.* 20-72 —The nature of the hydrate gel formed when cellulose material is mechanically processed in the presence of water depends on the chem. characteristics of the unprocessed material. A wood pulp of high α -cellulose content hydrates slower than ordinary sulfite pulp and yields a paper of higher resistance to tearing and folding. The strength of a sheet is ascribed to an interlocking of the fibers plus the cementing effect of the hydrate gel. Hydrated wood pulp of high α -cellulose content produces paper of quality about equal to paper made from the better grades of rag stock. F. A. SIMMONS

A contribution to the technic of fiber measurement. BRUNO SCHULZ. *Papier-Fabr.* 29, Tech.-Wiss. Teil 4-5(1931) —A simple projection microscope and a method of sampling for making slides are described. The prep'd fibers are collected on an 80-mesh or finer wire cloth, and a small portion is transferred to the slide with a needle. Lifting fibers directly from dil. suspension with a needle results in loss of fines, while if a puppet is used the sample is unduly enriched in short fibers. R. H. DORTCH

Chemistry of Australian timbers. W. L. CONYER AND H. E. DADSWELL. *J. Council Sci. Ind. Research* 4, No. 1, 45-8(1931) —A research project is outlined. For eucalypt Australian woods, the usual methods of wood analysis had to be modified because of their resistance to the action of Cl used in α -cellulose. These woods need a much longer treatment than usual and the cellulose appears to be greatly disintegrated. The repeated chlorinations cause the cellulose to become gelatinized by the subsequent alkali treatment and it cannot be filtered. The α -cellulose detn. is of value in detg. the suitability of the cellulose for the manuf. of rayon and cellulose lacquers. Chem. analysis is the only means of positively differentiating jarrah and karri. The former shows 38.8-52% cellulose while the latter yields 53.7-63.7%. It is hoped similarly to differentiate hoop pine and bunyah pine. Lignin results are often high because the Calif. mixt. which is used for extn. leaves behind some of the extraneous materials frequently present in woods. These substances contaminate the lignin and vitiate the results. C. R. FELLERS

The penetration of water vapor into wood. L. M. PIDGEON AND O. MAAS. *Research Notes* 3, 47-50(1930). *Pulp Paper Mag. Can.* 31, 530-3(1931). —See C. A. 24, 4155. A. PAPIREAU-COUTTIER

Wood-pulp evaluation for soda pulp manufacture. H. W. MORGAN. *Paper Trade J.* 92, No. 15, 51-2(1931) —The importance of careful control of the quality of soda pulpmaking and the present methods of grading are briefly discussed. The technic of the 1% alkali sol. test has been modified to give more uniform heating (resulting in slightly higher temp.), which gives more reproducible, though somewhat higher, results. From the results of evaluation of a given wood by phys. and chem. methods, the relation of the yield to the 1% alkali sol. material was expressed graphically. It is estd. that an increase of 1% in alkali sol. material is accompanied by a decrease of 3% by wt. in the pulp yield. The effects of decay on the compn. of wood are discussed and the relation of these effects to the pulp yield is pointed out. A. PAPIREAU-COUTTIER

The growth of molds on sulfite pulp, the consequent fiber degradation and cellulose decomposition. WALTER DRECHSEL. *Papier-Fabr.* 28, Tech.-Wiss. Teil 709-13 739-46, 845-54(1930). 29, 5-9(1931) —The method generally used in studying this problem is to follow the wt. changes on decompn., which fails to show the mechanism of fungus attack. The presence of nutrient media may also lead to incorrect conclusions. D. followed the decompn. microscopically, and later by the rate of evolution of CO₂. The behavior of 15 molds toward sterile bleached and unbleached sulfite pulp was studied, at 650 magnification, special technic being used. The results, which are copiously illustrated, showed that all the molds tried attacked both pulps, but in different ways. In certain cases, and at certain stages of decompn., the fibers appeared as if swelled in peptizing media, the end result in all cases was a structureless mass. Lignin and other impurities in unbleached pulps seemed to serve as nutrients. Humic substances were formed, especially from the more lignified pulps. It was found that the molds are not anaerobic, and cannot exist solely on the O content of the cellulose. It was possible to control conditions in the lab. so that the lignin decompn. was greater than the cellulose decompn., but the results hold no promise of a successful process for isolating pure and strong fibers by biological decompn. of lignocellulose. While most

pulp in storage may be attacked by molds if special care is not taken to prevent infection, the attack will be localized since in the absence of special nutrient media the molds cannot propagate greatly. Though the weight loss is small, the strength loss may be large, because of the nature of the attack on the fibers. A bibliography of 60 items is appended. R. H. DOUGHERTY

Description and historical development of the manufacture of sulfite spirit from waste pulp liquors in Germany during the war period 1914-18. WALTER SEMBRITZKI. *Papier-Fabr* 29, Tech.-Wiss. Teil 69-74, 85-92(1931).—Methods, operation and costs are reviewed. R. H. DOUGHERTY

Review of patents on the use of waste liquors and gases of the pulp industry, for the period 1925-30. W. SCHMID. *Papier-Fabr* 28, Tech.-Wiss. Teil 573-5, 587-90, 620-3, 636-9, 671-2, 727-9, 744-6, 875-7(1930), 29, Tech.-Wiss. Teil 10-4(1931). R. H. DOUGHERTY

The galvanic behavior of a chromium-nickel-iron alloy in sulfite liquors. W. ANDREW WESLEY and F. L. LAQUE. *Paper Trade J* 92, No 16, 56-61(1931).—Cr-Ni-Fe alloy can function as cathode in galvanic couples in neutral and in acid solutions without destruction of the surface film which renders it passive in these electrolytes. Mechl. and other disturbances of the surface of this alloy rendered it active, at least temporarily, in which condition it behaved more nearly like Fe than like a noble metal. It coupled with Cr-Ni-Fe in Na_2SO_3 soln. suffered just the same amt. of galvanic corrosion as it did when coupled with Pt. In sulfite liquors, freshly surfaced Cr-Ni-Fe behaved erratically for a time, but always tended to become passive, to behave like a noble metal and to remain in that condition until its surface was again abraded. The behavior of couples in which a valve bronze was connected with the alloy or with Pt indicated that the corrosion process was different in tower acid from that in reclaimed acid. The test conditions were somewhat different from plant conditions in that these solns. were satd. with SO_2 at 50° and atm. pressure. In reclaimed acid the corrosion of valve bronze was accelerated to the same degree when it was coupled with Cr-Ni-Fe as with Pt, indicating that the corrosion process involved to a predominating extent the reaction of some depolarizing agent active in this soln. In tower acid the corrosion of valve bronze was accelerated much less by coupling it with Fe alloy than by coupling it with Pt. A similar result was obtained in pure dil. H_2SO_4 soln. The discrepancy is thought to be the result of a difference between the H overvoltages of the 2 cathode surfaces. A. PAFINEAU COUTURE

"Liquid rosin" and its possible uses. W. SCHMID. *Papier-Fabr*, 29, Tech.-Wiss. Teil 1-4(1931).—Quality and compn. of the liquid rosin that separates from alk. pulp waste liquors varies with the process and the wood used. An av. yield is 30 kg. per ton of pulp made. Purification and uses of this material are described. The phytosterol, present up to about 2%, is of increasing importance. R. H. DOUGHERTY

Process water in the pulp and paper industry. A. SPITZGERBER. *Papier-Fabr*, 29, Tech.-Wiss. Teil 81-4(1931).—The impurities of water and their removal from boiler feed, washing and cooling water are described. R. H. DOUGHERTY

White water in paper and pulp mills and its utilization. R. J. MARX. *Trans Inst. Chem. Eng.*, Advance copy, Dec., 1930, 24-9.—White water is the effluent from the passage of the aq. suspension of fiber over a close meshed wire in a paper machine. Settling methods for fiber removal were used first, but these are mostly of low efficiency. An early power driven machine employed an endless belt traveling over a wire-covered drum. The Dorr save-all consists of a vertical tank with horizontal trays connected close to the vertical shaft, which carries one scraper in each chamber. The Oliver save-all is a horizontal revolving cylinder with a wire screen in which the deposited fiber forms a filtering medium. It has oscillating rakes and works under suction. M. prefers a gravity settler, in which the water is allowed to fall as drops on to a distributing surface. The impacts cause the sepn. of entrained air from the fiber, which thereafter settles readily in a large conical tank with bottom discharge for solids. B. C. A.

History of paper-products manufacture. H. RAFF. *Kunststoffe* 20, 272-5(1930). E. J. C.

Chemistry and paper making. M. BROU. *Chimie & Industrie* 25, 480-7(1931).—A general discussion of the importance of chemistry in the various pulp- and paper-making processes. A. PAFINEAU COUTURE

Copper-pipe advantages in the paper industry. CHARLES A. HILL and LELAND T. SUMMERS. *Paper Trade J* 95, No 16, 61-2(1931).—A brief discussion is given of the advantages of Cu over other metals for piping in pulp and paper mills. Its use has been rendered quite practical and economical by the development of stream line hard Cu.

pipe and fittings which is made threadless and designed to make a conveniently soldered connection to thin wall Cu pipe. A PAPINEAU-COUTURE

Preparation of paper-making materials. I. Rags, esparto, wood pulp. J. P. BARBOUR. *World's Paper Trade Rev* 95, 1114, 1120, 1129(1931).—A brief discussion, based on 60 yrs.' practical experience, of the prep'n of rags, esparto, straw and wood pulp, more particularly as regards boiling. A PAPINEAU-COUTURE

Methods of determining acidity in paper. P. F. WIMMER, et al. *Paper Trade J.* 92, No. 16, 63-4(1931).—Report of the Technical Assoc. of the Pulp and Paper Ind. sub-committee on paper testing chem. methods. Thirteen labs. have agreed to cooperate in an investigation carried out on 6 different grades of paper by 5 different methods (which are described) including 2 pH methods (electrometric and colorimetric). A P.-C.

The use of liquid chlorine in the preparation of bleach liquors. H. R. J. FEENEY. *World's Paper Trade Rev* 95, 1092-6, 1154, 1156(1931).—After a brief outline of the advantages of liquid Cl_2 its use for the chlorination of bleaching powder and of milk of lime in paper mills is briefly outlined. A PAPINEAU-COUTURE

The volumetric composition of paper. I. Permeability of paper to air. T. TREVOR POTTS. *World's Paper Trade Rev* 96, 1257-66, 1330(1931).—After a brief discussion of the primary requisites for permeability measurements, the effect of changing conditions of testing on the rate of flow of air through paper was studied. The rate of flow is proportional to the area of the sample, provided that it is not so great that, in papers of low permeability, it will be slightly modified at the pressure required to obtain flow of air through the paper. For all practical purposes, the flow is practically proportional to the pressure difference between the 2 sides of the sample. No change in rate of flow occurs with time. The rate of flow is proportional to some power of the thickness, the numerical value of the power varying according to the sample of paper tested. This is taken as *exptl.* proof that the inter fiber spaces of paper are not simple capillaries, and that the thickness of the sample should not enter into the expression of permeability until more is known about its meaning. A PAPINEAU-COUTURE

Can the bursting pressure for a [paper] test area of any size be calculated if it is known for one area? G. BIRKETT AND BARTO SCHULZ. *Wackm. Papierfabr.* 61, 1652-5(1930). *Papier Fabr.* 29, Tech.-Wiss. Teil 231-5(1931).—The relation $p/P = (f/F)^{1/x}$ was found to hold, the value of x lying between 1.6 and 1.8, with the value 1.7 giving satisfactory approx. results in all cases. In the formula, p and P are the bursting pressures for the areas f and F . In a perfect homogeneous material, x should equal 2. The formula fails for very weak papers because of the *exptl.* error in detg. small pressures. R. H. DOUGHERTY

Grease resistance of paper. ROGER C. GRIFFIN. *Paper Trade J.* 92, No. 15, 44-51(1931).—Results of a collaborative study of the turpentine penetration test and the red oil penetration test, applied to greaseproof paper and vegetable parchment, are reported. It is recommended that paper be first tested in a preliminary way by the turpentine penetration method using 4 sep. pieces, and that, if the *av.* result of this test is less than 5 mm., the paper be tested by the red oil penetration method. Adoption of both methods (the technic of which is described in detail) as tentative standard methods by the Technical Assoc. of the Pulp and Paper Industry is recommended. A method proposed by Ernest Scheller, which is based on the no. of transudations per unit there are in a definite period of time and the technic of which is described, will be studied. A PAPINEAU-COUTURE

Newsprint control and tests. RAYMOND FOURNIER. *Papeterie* 53, 350-61(1931).—A discussion is given of the advantages to be derived by newsprint users, particularly large dailies, from the control of the paper purchased, with an outline of the principal tests and the instruments required. A PAPINEAU-COUTURE

Manufacture of insulating board from cornstalks. O. R. SWEENEY and W. E. EMLEY. *U. S. Bur. Standards, Miscellaneous Pub.* No. 112, 27 pp (1930).—A description is given of the *exptl.* work at Iowa State College on the manuf. of cornstalk insulating boards, including lab. scale and *semicom.* *expts.* Both chem. and mech. methods were used. The boards produced appeared to have the same properties as boards from other sources. A statistical survey of the quantity, distribution, value, etc., of cornstalks is given and cost data are presented in pointing out the possibilities of cornstalks in board production. M. HERNIG

X-ray study of the gelatinization of nitrocellulose (DESMAROUX, MATHIEU) 2. Plastic materials (Fr. pat. 698,321) 18.

HESS, WALTER F.: *Die Praxis der Papier-verarbeitung*. Berlin: M. Krayn. 256 pp

Cellulose. AKTIESELSKAPET RAUL PICTET & F. THASALDSEN. Fr. 609,338, July 3, 1930 Cellulose is obtained by treating wood, etc., with a lye contg a much larger amt. of free SO_2 than in the usual "bisulfite" process, less lye being used than is necessary to recover entirely the primary material treated

Cellulose derivatives. I G FARBERND A-G. (Max Hagedorn and Eugen Gühning, inventors). Ger 516,882, Apr. 18, 1929 The derivs are recovered from their solns by pptn with steam or vapor Thus, a soln of cellulose laurate, in PhCl , pyridine and pyridine-HCl is treated with steam at 120° The cellulose laurate is pptd as a fine powder and can be scpd by centrifuge or filtration A soln of cellulose nitrate in MeOH is treated with CaH_2 vapor to ppt the cellulose nitrate Further examples are given

Alkylcellulose masses. I G FARBERND A-G (Gerhard Balle and Kurt Sponsel, inventors) Ger 516,751, Aug 14, 1925 Masses or objects are prepd partly or wholly of aq sol alkylcellulose by pressing layers or plates of this substance on one another with simultaneous treatment with water, moisture or aq solns

Benzylcellulose. I G FARBERND A-G (Eduard Dörr, inventor). Ger. 522,170, Aug 18, 1925 See U S. 1,771,529 (C A 24, 4631)

Cellulose acetates. I G FARBERND A-G. Fr 608,392, July 4, 1930 Primary acetates of cellulose sol in acetone are prepd by acetylating monohydroxyalkyl ethers of cellulose with Ac_2O , AcOH and a catalyst at a temp below that at which hydrolysis takes place

Cellulose esters. KODAK PATHE Fr 608 689, Oct. 7, 1929 In making cellulose esters the diluting reaction acid is replaced wholly or in part by a solvent for primary esters of cellulose, particularly CH_2Cl_2 A reflux app is used

Treatment of cellulose esters. CAMILLE DREYFUS and CLIFFORD I HANEY (to Camille Dreyfus) Can 309,756, Mar 24, 1931 Cellulose acetate that has been prepd by the reaction of cellulose with Ac_2O in the presence of H_2SO_4 is purified by sepg it from the primary soln. in which it is formed, dissolving in AcOH and then adding $\text{Ba}(\text{OAc})_2$

Cellulose ethers. I. G FARBERND A-G (Gerhard Balle and Karl Ost, inventors) Ger. 522,054 Feb 11, 1928 Cellulose ethers insensitive to water are prepd by treating metal compds of cellulose with an excess of an alkylating agent contg up to 2 mols of an alkyl halide for each 10-30 mols of alkylating agent The reaction is effected under such conditions of temp and pressure that sapon of the excess of alkylating agent is avoided Thus, an alkali cellulose prepd from 1 mol of cellulose ($\text{C}_{12}\text{H}_{10}\text{O}_2$) may be heated with EtCl 14 and PhCH_2Cl 1 mol to 110 - 120° in a pressure vessel Cf C A 25, 1379

Carbohydrate ethers. I G FARBERND A-G (Otto Leuchs and Eduard Dörr, inventors). Ger. 522,171, Nov. 7, 1928. See Fr 684,330 (C A. 24, 5496)

Nitrocellulose. FREDRICK OLSEN. U S 1,798,270, March 31, Cellulose is nitrated, pulped and then alternately boiled and washed in water. in order to effect purification.

Dissolving and dispersing nitrocellulose. WILLIAM B. PRATT and RALPH T. HALSTEAD (to Dispersions Process Inc) Can 311,079, May 5, 1931 Cellulose is dissolved in wood creosote and a hydrophilic colloid incorporated therein and water is gradually added until a change of phase takes place and the nitrocellulose soln disperses as minute particles in the aq medium

Use of cellulose compounds such as celluloid solutions for sealing joints between watch crystals and bezels. WARREN F. BLECKER. U S 1,799,146, April 7.

Films. KODAK-PATHE Fr 698,727, Oct. 12, 1929 In making films from cellulose derivs the sheet of collodion is submitted, during the formation of the film, and particularly before solidification, to a drawing in the direction of its breadth to avoid the formation of wrinkles or creases The drawing may be by bands of fabric on the edges of the sheet which they accompany into the bath guiding it to the roller

Artificial silk. HÖLKENSEIDE G M B H Ger 522,469, Nov. 17, 1928 See Fr 684,658 (C. A. 24, 5497).

Artificial silk. HENRY DREYFUS. Fr. 697,427, June 14, 1930. Artificial threads or filaments are made by extruding a spinning soln of cellulose acetate or other cellulose deriv. into an evaporative medium in which an external skin-like layer is formed on the threads The threads are passed into an atm of solvent vapor in which the external

layer is softened, and then drawn. The drawing is limited to the part of the filaments which have been submitted to the softening. An app. is described.

Artificial silk. HENRY DREYFUS. Fr 697,725, June 20, 1930. A soln of cellulose or its deriv. is extruded and the threads are wound or twisted in a centrifugal box of relatively small diam and at a very high speed, *e.g.*, 12 000 to 20 000 r.p.m.

Artificial silk. CAMILLE DREYFUS. Fr 697,814, June 23, 1930. A soln of an org. or inorg. deriv. of cellulose such as cellulose acetate in a solvent is extruded through orifices at the outlet of which the soln. is received into a coagulating bath, the nature of the solvent being such or the coagulation being carried out in such a way that a certain amt. of the solvent remains. The threads are then heated and are obtained in a transparent brilliant state. Thus, threads made from a soln. of cellulose acetate in ethylene dichloride, MeOH and inacetin are passed into a coagulating bath contg. water and then over a heated roller. Other examples are given.

Artificial silk. CARL HAMEL A-G and EDMUND HAMEL. Fr. 697,717, June 20, 1930. App. for regulating the course of the guide thread is described.

Artificial silk. JACQUES DELPECH and CONSTANTIN HEINRICH. Fr 699,423, July 5, 1930. An artificial silk of very fine strand is made by partially dehydrating the nitrocellulose by means of 70% MeOH, dissolving the partially dehydrated nitrocellulose in MeOH, and spinning this colloidal into 10-40% MeOH at a temp. not below 25°.

Artificial silk. COMPTOIR DES TEXTILES ARTIFICIELS (SOC. ANON.). Fr 697,853, June 24, 1930. In making artificial silk, the thread after passing through the coagulating bath is passed through a narrow tube traversed by a liquid. The liquid is caused to flow in the opposite direction to the movement of the thread to increase the tension or with the thread for the reverse effect. The liquid may be the same as that in the bath or may be a washing liquid.

Artificial silk. COMPTOIR DES TEXTILES ARTIFICIELS (SOC. ANON.). Fr 698,660, July 8, 1930. A brilliant pearly artificial silk is made by transforming a cellulose rich in α -cellulose into a viscose rich in cellulose, preferably more than 7.5%, and, after a short ripening, spinning the viscose in a bath of H_2SO_4 and Na_2SO_4 , the content of Na_2SO_4 being more than double that of H_2SO_4 and the latter being 100-150 g. per l.

Artificial silk. I G FARBERNIND A-G. Fr 697,329, June 13, 1930. Hollow threads of artificial silk are made by dissolving cellulose derivs. such as cellulose acetate in one or more solvents adding one or more non solvents, the b.p. of which is higher than that of the solvents, heating the spinning soln. and spinning into a gaseous medium, the temp. of which is lower than the b.p. of the non solvents. Examples are given in which cellulose acetate is used in a mixt. of acetone and toluene or CH_2Cl_2 .

Artificial silk from viscose. I G FARBERNIND A-G. Fr. 697,997, Mar. 24, 1930. Uniform solns. of viscose are prepd., the use of compressed air or other gas being completely avoided. The filtration, elimination of air and the final mixing of the distinct contents of the dissolving reservoirs are effected without any interruption. The crude soln. to be spun coming from the reservoirs is driven back by a common pump to the filter presses connected to a common distributing pipe, then the filtrates from all the filter presses working simultaneously are collected in a common conduit.

Artificial silk and films. I G FARBERNIND A-G. Fr 698,100, June 26, 1930. The reaction mixt. of cellulose esters obtained in liquid SO_2 by Fr 664,459 (C.A. 24, 961) is used, without isolation of the dissolved cellulose esters, for making threads or films by dry spinning or flowing the liquid into a pptg. liquid.

Artificial silk, films, etc. WOLFF ET CO. KOMMANDIT GES. AUF AKTIEN, EMIL CZAPEK and RICHTARD WEINGAND. Fr 698,307, July 2, 1930. In the manuf. of threads, sheets, etc., from aq. solns. of cellulose and a coagulating bath, the initial soln. of cellulose is heated to about 50°.

Artificial silk, films, etc., of reduced luster. HENRY DREYFUS. Fr. 698,093, June 26, 1930. TiO_2 , preferably in a finely divided form, is incorporated in the solns. of org. derivs. of cellulose used for making silk, films, etc. The size of the particles of TiO_2 is preferably 0.0001-0.00035 mm. and dispersing agents for the particles may be used.

Artificial silk, plastic materials, etc. OSCAR KOHORN & Co and HELLMUTH SCHTUFF. Fr 697,471, June 16, 1930. Solns., lacquers, films, artificial silk, etc., are prepd. by treating cellulose acetobutyrate with mixts. of liquids which have not marked solvent properties for this substance, *e.g.*, mixts. of alcs. with hydrocarbons or with their derivs. Examples are given using mixts. of EtOH and C_6H_6 in equal amts.

Artificial silk threads, etc. CAMILLE DREYFUS. Fr 697,573, June 17, 1930. In dry spinning artificial threads, etc., in an upward direction, the operation is started by surrounding the nozzles with a coagulating liquid such as water, C_6H_6 , MeOH, EtOH or CCl_4 , which is run off once the threads have begun to form.

Artificial fibers. I G. FARBENIND A-G (Adolf Kämpf, inventor). Ger. 522,368, May 1, 1929. A method of stretching dry-spun artificial fibers is described. The fibers are treated, in the course of the process, with swelling agents.

Preparing artificial silk fiber for spinning. MEINRAD F. THOMA. U. S. 1,799,399, April 7. The fiber is treated at a moderately warm temp. in a bath of water contg. an oil soap with a trace of H_2SO_4 , glycerol, NaOH and pinol.

Spinning and stretching artificial silk. FR. KÖTTNER A-G. Ger. 516,789, May 27, 1928.

Spinning apparatus for artificial silk. MARTIN HÖLKEN G. M. B. H. Ger. 516,788, July 31, 1929. Details of supplying the soln. to be spun to the spinning heads are given.

Nozzle for spinning artificial silk in acid spinning baths. ERSTE ÖSTERREICHISCHE GLANZSTOFF FABRIK A. G. Austrian 121,970, Nov. 15, 1930.

Cellulose acetate yarn. WILLIAM WHITEHEAD and CAMILLE DREYFUS (to Camille Dreyfus). Can. 310,732, Apr. 21, 1931. An acetone-sol cellulose acetate having an acetyl value of 52.5–56% is dried until it contains less than 1% of water, dissolved in substantially pure acetone, filtered and dry-spun at a temp. of 50–55°, whereby yarns are formed that are highly resistant to the delustering action of boiling water, and which have a strength that is 15–25% greater than that of yarn produced from a cellulose acetate soln. in 95% acetone.

Cooking fibrous material. THOMAS L. DUNBAR. U. S. 1,798,523, March 31. In a process such as digestion with acid liquor, gases and vapors from a digester are discharged through a conduit into an enlarged mixing chamber, and cold acid liquor is constantly fed into the chamber and the gases and vapors are thus utilized to preheat and pre-condition the liquor, this operation is conducted and the liquor is withdrawn from the chamber under substantially atm. pressure and the liquor is then forced into an accumulator maintained at superatm. pressure. An arrangement of app. is described.

Sulfite liquor. THOMAS L. DUNBAR (to Chemipulp Process, Inc.). Can. 309,887, Mar. 31, 1931. A process and an app. are disclosed for absorbing SO_2 gas in an soln. to obtain a sulfite liquor. Hot, strong H_2SO_4 soln. is mixed with cooler, dil. H_2SO_4 soln. in the presence of solid lime-contg. material. The resulting soln. is cooled, SO_2 gas is absorbed therein, in the presence of solid lime-contg. material, and unabsorbed SO_2 is passed from the last-mentioned absorption into contact with water in the presence of solid lime-contg. material to prepare a dil. soln. to be mixed with further quantities of hot strong H_2SO_4 .

Washing sulfite pulp. SYOSHI OKUDA. Japan. 90,207, Feb. 9, 1931. Stuff from the pulp separator in the manuf. of ordinary sulfite pulp sometimes contains acid, which lowers the quality of the pulp. In this process the stuff is mixed with milk of lime in a flow-box; by this means the acid in the stuff is neutralized and the quality of the pulp is improved.

Cellulose pulp. GEORGE A. RICHTER (to The Brown Co.). Can. 309,868, Mar. 31, 1931. Wood pulp, previously liberated by the action of an alk. liquor, is digested in an acid sulfite liquor to remove components other than α -cellulose, including ligneous matter and pentosans; then the pulp is digested in an alk. liquor to remove more non-cellulose components, and finally bleached.

Chemical wood pulp. GEORGE A. RICHTER (to The Brown Co.). Can. 309,867, Mar. 31, 1931. Raw cellulosic material is cooked in ammoniated water to dissolve a portion of its cementitious content and then cooked in a liquor contg. both NaOH and Na_2S to effect fiber liberation.

Pulp for paper making. JOHN D. RUE (to Champion Fibre Co.). U. S. 1,798,987, March 31. An unbleached and impure chem. wood pulp is partially bleached and washed, thickened, admixed with an alk. treating soln. such as caustic alkali and digested. Treating pulp is sepd., a portion of the sepd. soln. is coned., and the remainder of the sepd. soln. is combined with a portion of the so-coned. soln. The resulting alk. soln. is reused with added caustic alkali as alk. treating soln., and the remainder of the so-coned. soln. is used in alk. digestion of wood.

Bleaching paper pulp. JOHN NEUMANN. U. S. 1,799,601, April 7. The pulp is bleached in a bath contg. salt and oxalic acid and is then washed.

Paper. CHARLES J. BARNES. U. S. 1,799,350, April 7. Mech. features.

Paper containing rubber. GENERAL RUBBER CO. Fr. 697,683, June 19, 1930. A protective colloid such as starch esters known as "feculoid" or "feculose" or a protein is mixed with paper pulp. The mixt. is rendered alk. and a dispersion of rubber added. The pulp obtained is then treated in the usual manner for making paper. Cf. C. A. 24, 5159.

- Paper-making apparatus. HERMAN L. KUTTER (to Black Clawson Co.) U S 1,708,887, March 31 Structural features
- Fourdrinier paper-making apparatus. SAMUEL MILNE U S 1,709,801, March 31 Structural features
- Paper-making apparatus having a raising felt. EUGENE M. VERAON (to Papeteries Navarre) U S 1,708,821, March 31 Structural features
- Suction roll for paper-making apparatus. CARL F. BEARY (to Beloit Iron Works). U S 1,709,775, April 7 Structural features
- Suction box for paper-making, etc., machines. ERNST KRAMER Austrian 121,848, Nov 15, 1930
- Drying continuous paper sheets or other moist materials with air currents. JOHN P. BROWN U S 1,709,718, March 31 App and various details of operation are described

Apparatus for dewatering paper pulp, etc. RICHARD KASTNER and HERBERT SCHMOLKA Austrian 121,709, Nov 15, 1930 Addn to 115,901 (C A 24, 2295) and 120,690 (C A 25, 1996)

Apparatus for dewatering cardboard, etc., by compression. CARL WEISHAN Austrian 121,801, Nov 15, 1930

Sizing paper. JUDSON A. DECEW (to Process Engineers, Inc.). U S 1,709,216, April 7 A stream of sizing soln is mixed with a stream of alum and the mixed materials are then mixed with a stream of pulp stock U S 1,740,217 describes incorporating emulsified rubber particles into cellulose paper stock by mixing the rubber emulsion with the paper stock after the stock has been discharged from the beater, and adding and mixing a coagulant such as $Al_2(SO_4)_3$ to fix the rubber on the cellulose fibers, then forming the latter into paper Cf C A 24, 2004

Safety paper. RAYMOND E. BOHRER (to Todd Co., Inc.) U S 1,709,499, April 7. A plurality of max. valuation legends are printed in a concealed manner on documents such as bank checks, by use of substances such as "colorless inks" which are altered by chemical reagents such as "ink eradicators" so as to render the legends visible

Packing, etc., paper. FELIX THUNERT. Ger 519,867, Aug 20, 1923 Addn to 518,099 (C A 25, 2568) In the method of Ger 518,099, ornamental effects are obtained by spraying colored or uncolored adhesives on to the cellulose wadding

Wallboard and pressboard, etc. FERNANDO S. VIVAS. U S 1,800,121, April 7. In producing a sheet of felted cellulosic fibers, the pulp used is beaten with a soln of dextrin

Carbon paper. FIRMA GUNTHER WAGNER Ger 516,975, Mar 8, 1928 Carbon paper is prepd by giving the paper a layer of natural or artificial rubber, gutta serena, balata, dye and optionally a softening agent such as triethyl phosphate, oil, fat, fatty acid, etc., and drying Vulcanizing agents may also be added Examples are given Cf C A 24, 3903

Paper colored with sulfonic acid compounds of dyes obtainable from higher molecular diarylaminoquinones. GEORG KRÄNZLEIN, HEINRICH GRÜNER, MAX THIELER and KARL HAGER (to General Aniline Works) U S 1,708,544, March 31 Dyeings of good fastness to water and to light are effected by coloring paper with a salt of a sulfonic acid of a compd of the probable general formula



wherein C_1 and C_2 , on the one hand, and C_3 and C_4 , on the other hand, represent C-atoms belonging to a naphthalene or carbazole nucleus, or to a benzene nucleus which is substituted by at least one substituent of the group consisting of PhO , $PhNH$, $BzNH$, alkylamino, alkyl and halogen Examples with details of procedure are given, involving the use of the Na sulfonate salts of the dyes produced from (1) 2,5-di(*N*-ethylcarbazolyl-3'-amino) 3,6-dichloro-1,4-benzoquinone, (2) di(*β*-naphthylamino)dichlorobenzoquinone, (3) 2,5-di(*o*-aminodiphenoxy)-3,6-dichloro-1,4-benzoquinone, and, (4) 2,5-di(*p*-aminodiphenylamino) 3,6-dichloro-1,4-benzoquinone Blue, violet, red and bluish gray tints may be obtained

24—EXPLOSIVES AND EXPLOSIONS

CHARLES E. MUNROE AND C. G. STORM

Recovery of ammonium nitrate from amatol. Explosion at ammonium nitrate plant of Ammonite Company. R. NORRIS SHREVE *Ind Eng Chem* 23, 566-73(1931).—S. describes in detail the process and equipment used in the Ammonite Co's plant for recovery of NH_4NO_3 from water soln obtained from amatol (TNT NH_4NO_3 mixt.) salvaged from loaded shell. He also discusses the effects and probable causes of an explosion which destroyed the plant and killed 16 persons on March 1, 1924.

C. G. STORM

Effect of impurities on the freezing point of nitroglycerin. A. SAPOZHNIKOV AND K. SVITKO *Zhur Prikladnoi Khim* 3, 1077-85(1930).—Fps of the 2 isomers of nitroglycerin are 12.2° and 13.15° , resp. The use of NaNO_3 yields the high f.p., and the use of KNO_3 the low f.p. isomer. Addn of dinitroglycerol for lowering the f.p. is practical while addn of dinitrochlorohydrin or tetranitrodiglycerin has only a small effect. Fps of these mixts contg 60-100% nitroglycerin are given.

V. K.

Apparatus for determination of flame propagation. J. ZAST *J usines gaz* 55, 57-8(1931).—The principle of the app. described is variation of height of combustion cone in a calibrated Bunsen burner. Curves and numerical examples of detns. are given.

B. J. C. VAN DER HOEVEN

Flame speeds in the "inflammation" and "detonation" of moist carbon monoxide-oxygen mixtures. WILLIAM A. BOVE AND REGINALD P. FRASER *Proc Roy. Soc (London)* A130, 542-51(1931).—The "detonation range" of moist satd CO-O mixts at about 18° and atm. pressure lies between about 40 and 80% CO contents, as compared with about 15.4 and 94.0% for their inflammation range. Within the said range the observed rates of detonation in a tube 1.3 cm. in internal diam. all lie between 1700 and 1800 m. per sec. There was a well marked max. rate of detonation at the 75CO/25O₂ mixt. compn.

H. C. PARISH

Application of the antioxygenic effect to the problem of extinguishing fires. Negative catalysis of the ignition of carbon. CHARLES DUPRAISSE AND RAYMOND HORCLOIS. *Compt rend* 192, 564-6(1931).—Ignition, like oxidation at low temps., may be hindered by neg. catalysts, and there is thus no theoretical reason why fires may not be extinguished by applying this principle. CCl_4 added to air so as to form 5% by vol., exerts this antioxygenic catalytic effect, extinguishing a wood charcoal fire in an expl. stove very quickly. That the extinguishing is not due to the massive action of the CCl_4 or to rarefaction of O₂ by diln. is shown by similar expts. in which the diluent is N₂ (scarcely any extinguishing effect) or where the air and CCl_4 are enriched with up to 31% O₂ (marked extinguishing effect). For CCl_4 the threshold of activity is at $1/10$ of satn. or $1/100$ of vol. CO₂ behaves similarly but not so markedly. Other substances acting like CCl_4 are CHCl_3 , $\text{C}_2\text{H}_5\text{Cl}$, $\text{C}_2\text{H}_5\text{Br}$, $\text{C}_2\text{H}_5\text{I}$, EtBr , ethyl chloroformate, sulfur chloride, thionyl chloride, sulfuryl chloride, PCl_5 , PBr_5 , As chloride, SiCl_4 , TiCl_4 , SnCl_4 , BCl_3 , Et_3NH , SO_2 , etc. POCl_3 is very active, its threshold being $1/20$ of satn. or $1/2000$ of vol.

ALBERT THOMAS FELLOWS

Prevention of mine explosions by means of water sprays. Wm. KOHLSCHIEIN. *Z ges. Schiess Sprengstoffw* 26, 77-81(1931).—A method is described for extinguishing an explosion flame by means of an atomized spray of water, automatically released by means of compressed air from 45 l. of water in a cylinder placed in the mine gallery about 20 ft. from the shot. The release is effected electrically by the current which fires the shot. Tests in a steel gallery at Freiberg, Saxony, in which 200-g. charges of dynamite were fired from a mortar, showed a length of flame of 25-30 m. without the water spray, and only 0.5-1.5 m. when the spray was used. To be effective, the spray must be released 2.5-3 sec. before the shot is fired, the delay of the shot being obtained by "delay detonators."

C. G. STORM

Prevention of explosions in mirror-silvering materials. MAX ERMES *Diamant* 51, 62-3(1929). *Ceram Abstracts* 9, 24.—In the prepn. of a silvering mixt. there is a possibility of the formation of CNOAg on the addn. of alc. to a concd. soln. of Ag_2O in NH_3 . Two kinds of silver fulminate, Howard's and Berthelot's, are formed and they are both more explosive than the best gunpowders. Precautions to be taken to prevent accidental explosions while prepng. and working with silvering solns. in mirror manuf. are discussed.

G. G.

Danger of explosion in silvering solutions. MAX ERMES *Diamant* 51, 587(1929). *Ceram Abstracts* 9, 725.—The statements of E. Lohmann (cf. following ab-

(tract) in regard to safety from danger of explosions in the Rochelle salt process are confirmed by 1. The mirrors of mirrors prep'd by this process that are still sound though they were prep'd 20 yrs. ago. Experience with a soln. that had been stored for a year confirmed the fact that the Brashear process is also safe if the solutions are properly dild. G G

Source of danger of explosion in chemical silvering of glass. EDUARD LOHMANN, *Diamant* 51, 526-30(1929). *Ceram. Abstracts* 9, 725—The danger of explosion is said to lie in the prep'n. of the ammoniacal soln. of Ag for the Brashear process, when soln. "A" is made alk. with KOH or NaOH. In certain concns the compd $\text{Ag}(\text{NH}_3)_2\text{OH}$ is formed, which breaks down to give AgNH_2 and Ag_3N . This does not happen in dil solns contg less than 0.35 g-atoms of Ag per l. The mixt. of AgNH_2 and Ag_3N , the "Knallsilber" of Berthelot, detonates violently on the slightest disturbance of the soln. In the Rochelle-salt process, where addns. of KOH are unnecessary, there is very little danger of the formation of this explosive compd, and the danger in the Brashear process can be minimized to a large extent by avoiding too concd solns. of Ag, by using dil caustic alkalis and by mixing the soln. only just before use. It is also good practice to keep the flasks which contain such solns. away from direct exposure to sunlight. G G

Explosions in production of mirror coatings. WERNER MULLER, *Diamant* 51, 42-4(1929). *Ceram. Abstracts* 9, 21—The chemistry of prep'g silvering solns. for mirror silvering is discussed, and it is pointed out that there is a possibility of formation of such explosive compds as AgNH_2 by the combination of Ag and part of the NH_3 , or in case other ingredients are used in prep'g the solns., other explosive compds. may be formed. G G

Explosion studies of ammonia-air and ammonia-oxygen mixtures (with regard to high pressures). H. HERRMANN FRANCK AND GERHARD DÖRNING, *Z. anorg. Chem.* 44, 273-7(1931)—An investigation of the lower explosion limits of mixts. of NH_3 with air and O was carried out in a 50-l. bomb at an initial pressure of 1-20 atm., with various means of ignition. Independent of the nature of the means of ignition, there is a sharply defined explosive limit above which explosion is propagated without interruption in all directions. With less efficient ignition (elec. sparks or small quantities of combustible mixts.), complete reaction results only with increased concn. of NH_3 in the gas mixt. With more efficient ignition (e. g., 0.5-1 g. KClO_3 , C mixt.), the explosion limits were, for 1 atm., 15.5% NH_3 with air, and 13.5% NH_3 with O, for 20 atm., the limit fell to 10% NH_3 with O. C. G. STORM

Explosions in mixtures of acetylene and electrolytic gas. W. A. BOVEY, R. P. FRASER AND F. LAKE, *Proc. Roy. Soc. (London)* A131, 1-17(1931), *J. C. A.* 21, 2064—Explosions of mixts. of $x(\text{C}_2\text{H}_2) + (100-x)(\text{O}_2 + 2\text{H}_2)$ were studied by the open tube and the closed tube methods. The flame speeds, S , were measured and the flames photographed. Initial S is reduced on addn. of 1% C_2H_2 , further addns. up to 11.8% C_2H_2 greatly increase S which reaches a max. at $\text{C}_2\text{H}_2 + 6\text{O}_2 + 10\text{H}_2$, when the O_2 is sufficient to burn C_2H_2 to $\text{CO}_2 + \text{H}_2\text{O}$, increasing the C_2H_2 content above 11.8% causes a rapid decrease in initial and general S until a 40 $\text{C}_2\text{H}_2/60(\text{O}_2 + \text{H}_2)$ mixt. is reached. No C is deposited in mixts. contg. 25% or less C_2H_2 , i. e., in mixts. with C_2H_2 less than $\text{C}_2\text{H}_2 + \text{O}_2 + 2\text{H}_2$, the combustion is given by $\text{C}_2\text{H}_2 + \text{O}_2 + 2\text{H}_2 = 2\text{CO} + 3\text{H}_2$. Data with A has no effect, while data with N_2 causes C depn. and steam formation. In no expts. with explosions of $x(\text{C}_2\text{H}_2) + (100-x)(\text{O}_2 + 2\text{H}_2)$ or $x(\text{C}_2\text{H}_2) + (100-x)(\text{O}_2 + 2\text{CO})$ was there any sep'n. of C or survival of C_2H_2 until x was greater than 25. With low C_2H_2 content the "water-gas" equl. was not attained during cooling behind the flame front. S. LENTHER

Smokes and smoke gases (SMOLCZYK) 13. The effect of an electric field on flames and their propagation (LEWIS) 2.

Ballistic powders. E. I. DU PONT DE NEMOURS & Co. Fr 698 308, July 2, 1930. A ballistic powder which may be dried contains nitrocellulose in combination with a polynitrotoluene and an ester of an org. acid, each of which has in the liquid state a marked solvent action on nitrocellulose and each of which is non-volatile, non hygroscopic and insol. in water, e. g., nitrocellulose 85, dinitrotoluene 10, dibutyl phthalate 7 and diphenylamine 1 part.

25—DYES AND TEXTILE CHEMISTRY

L. A. OLNEY

Recent developments in dyes and dyeing. CHAS. E. MULLIN AND ALFRED R. MACORMAC. *Textile Colorist* 52, 739-43, 815-8, 849(1930).—A review of the present trends in dyeing and latest color cards, with a discussion of the properties of the recently announced dyestuffs. CHAS. E. MULLIN

Action of light on some vat colors. W. F. A. ERMEN AND E. H. GOODYEAR. *Textile Manufacturer* 55, 437-8(1929).—Some abnormal effects in the fading of vat colors by light. CHAS. E. MULLIN

Dyeing and finishing costume cloths. H. JENNISON. *Textile Manufacturer* 57, 69(1931).—A general address on the dyeing and finishing of wool materials. C. E. M.

Notes on the dyeing of union fabrics containing acetate rayon. ENCO Russia 6, 489, 491(1931).—Brief practical notes are given on the production of various effects on such fabrics. A. PAPIEVEAU-COUTURE

Effect of valences of electrolytes in mordant dyeing. P. P. VIKTOROV AND A. A. LEBENEV. *Zhur Prikladnoi Khim* 3, 1175-88(1930). cf. C. A. 25, 1386.—For coagulating Benzo pure blue the effectiveness of electrolytes is as follows: $\text{NH}_4 < \text{Na} < \text{K} < \text{Mg} < \text{Ni} < \text{Mn} < \text{Zn} < \text{Cd} < \text{Al}$. V. KALICHINSKY

The resistance of coloring substances to ultra-violet light. III. GIUSEPPE A. BRAVO. *Boll. ufficiale staz. sper. ind. pell. mat. concanti* 8, 292-8(1930), cf. C. A. 24, 2888, 25, 2295.—Earlier investigations were continued. The dyes examined in this study belonged to the following classes: azo, pyrazolone, diphenylmethane, triphenyl methane, xanthene, acridine, quinoline, azine, oxazine and thiazine dyes. G. S.

Chemical control in the textile industry. W. F. A. ERMEN AND S. H. JENKINS. *Textile Manufacturer* 55, 279-81, 316-7, 352-3, 402-3, 439-40(1929), 56, 28-9, 71-2, 225-6, 261-2, 377(1930), 57, 70-1(1931).—A discussion of lab layout, the testing of coal and lubricants, the identification of synthetic yarns and starches, and the chemical analysis of flue gas, water, fabrics, sizing, finishing and weighting materials. CHAS. E. MULLIN

Use of the microscope in the textile industry. J. M. PRESTON. *Textile Manufacturer* 55, 293, 335-6, 379-80(1929), 56, 12-3, 130-1, 244-6, 358-9, 395-6(1930), 57, 6-7(1931).—Many excellent details of technique are given. C. E. M.

Some chemical differences between abaca and Canton fibers. HARTLEY E. SHERMAN. *Philippine J. Agr.* 1, 123-34(1930).—Color of the ash is of no value in differentiating abaca from Canton. Since both fibers are colored by $\text{K}_2\text{Cr}_2\text{O}_7$ soln this test is of no value in classifying Canton and abaca. Canton fibers were lighter in wt., showed lower tensile strength, lower elasticity, higher natural acidity, 2-3 times greater fat content and higher ash than abaca fibers of similar stripping. The "Mercerization Curl" produced by hanging the fiber in a tall tube of 20% NaOH gave tighter, linkier curls and greater contraction with Canton than with abaca fiber of similar stripping. Loss in wt. due to boiling the fiber for 1 hr. in 1% KOH soln was 6% higher for Canton than for abaca fiber, indicating a larger amt. of saponifiable matter and pectocelluloses in the Canton. JOHN O. HARDESTY

Relation between reeling process of the cocoon and physicochemical properties of the aqueous colloidal solution of sericin. II. KANEKO AND M. MIYASAKA. *Bull. Sericulture and Silk Ind. (Japan)* 3, 2-3(1930).—The sericin of silk cocoons has been examined as to its soly. in hot water and the stability, viscosity, surface tension, osmotic pressure, *sp. elec. cond.* and *degree of turbidity of its colloidal solns.* Increased ease of reeling appears to be correlated with lower soly. in hot water, higher stability of colloidal solns., lower surface tension, greater cond., greater osmotic pressure and greater turbidity. B. C. A.

Acetate silk developments. CHAS. E. MULLIN. *Textile World* 78, 1519, 1578(1930).—A general review of the developments in the use, dyeing, printing and processing of the acetate silks. CHAS. E. MULLIN

Recent developments in the delustering of rayon materials. A. J. HALL. *Textile Exporter* 6, No. 49, 14-5(1931).—A review of the recent patents. CHAS. E. MULLIN

What are the future trends in the American synthetic yarn industry? CHARLES E. MULLIN. *Chem. Markets* 28, 363-7, 492-6(1931). B. H.

Immunized cotton. GIOVANNI TAGLIANI. *Bull. soc. ind. Rouen* 58, 418-26(1930).—An address on the Zimmerman process of rendering cotton immune to the action of substantive dyes by treatment with toluenesulfonyl chlorides (cf. Ger. Pat. 396,926 (1922)). The chem. and phys. factors affecting the process are discussed.

in general terms. Among the properties of cotton so treated are the following: larger, stiffer, and more transparent fibers, maintenance of immunization against the action of dil acids, alkalis and bleaching liquors, also to the mercerization process if carried out rapidly, resistance to drying, storage and molds. The product is immune to most substantive dyes. It can be dyed with properly chosen S and vat dyes in which case certain phenolic materials and colloids are desirable adds to the bath. It also absorbs basic dyes, some azo dyes and several other classes. A list of patents on this type of process is given. I. H. ODELL.

Woolen and worsted cloth finishing. B. H. Textile Manufacturer 55, 314 A, 350 1, 442-3 (1929). — Carbonizing piece goods by the $MgCl_2$, $AlCl_3$, H_2SO_4 , and HCl processes, filling and weighting are considered. CHAS. E. MULLIN.

Sizing of worsted and woolen yarns. H. R. D. SHARP. Textile Manufacturer 56, 240-1 (1930). — General. CHAS. E. MULLIN.

Treatment and disposal of wool-washing effluent. BASIL A. SMITH. Textile Manufacturer 56, 452, 458 (1930). Trans Inst Chem Eng., Advance copy, Dec 1930, 16-23. — The acid-cracking solvent extn., and centrifugal processes of grease recovery, the recovery of scouring liquors by purification and the disposal of the final liquors are discussed. CHAS. E. MULLIN.

Treatment of suint liquors from wool-scouring. A. T. KING. Textile Manufacturer 56, 458-460 (1930). Trans Inst Chem Eng., Advance copy, Dec 1930, p 15. — The Duhamel and other systems are discussed. CHAS. E. MULLIN.

The Marwa washing process. M. DITTMER. Seifenfabr Zig 58, 235-7 (1931). — The Marwa process saponifies fatty acids with Na_2CO_3 in the washing machine in the presence of the wash goods. The evolution of CO_2 may aid cleaning but the odor of the fatty acids, the usual presence of neutral fat and traces of metals in the fatty acids are decided disadvantages and do not avoid the formation of Ca and Mg soaps as a result of water hardness. I. ESCHER.

Dyeing of furs (DOROGNEI) 29. Examination of textile oils (TRUESLER) 27. Combinations among certain dyestuff radicals (REBER) 10. Triphenylmethane dyes derived from quinoline, tetrahydroquinoline, diphenylamine and carbazole (SEN, SEN) 10. Dyes derived from oxalylidibenzyl ketone (SALFORS, CHAKRAVARTI) 10. Naphthalene-series. III. Action of bisulfite on *p*-nitrobenzenazo-5-naphthol (KOROTZKOVA, BILLOV) 10. Compositions of silicates with other compounds [dyes] (FR. PAT. 698,543) 17. Unflammable solvent containing hydrocarbons [for cleaning] (FR. PAT. 697,500) 13. Polymerizing oils and fats [for textile industry] (FR. PAT. 697,755) 27.

KRÄNZLIN, G. AND MARCUS, A. Baumwolle. Leipzig. Deutscher Auslandsverlag. 160 pp. Linen, M 6.

SCHULTZ, GUSTAV. Farbstoffstabellen. 7th ed. Band I. Lfg 16-17. Leipzig. Akad. Verlagsges. Pp 673-764. M 16. Cf C A 25, 2855.

TOBLEK, FRIEDRICH. Sisal und andere Agavefasern. Leipzig. Deutscher Auslandsverlag. 104 pp. Linen, M 5.

WALLAND, HEINRICH. Einführung in die quantitativen textiltchemischen Untersuchungen. Vienna. Holder Pichler Tempky A-G. 206 pp.

Dyes. OSTRO RESEARCH LABORATORIES, INC. FR. 698,076, June 24, 1930. 4 (*p*-alkoxyphenyl)azo-*m*-phenylenediamines in which 1, 2 or 3 alkyl groups may be substituted in each benzene ring are prep'd by diazotizing an alkoxyaniline and coupling with a diaminobenzene. Thus *p*-phenetidine is diazotized and coupled with *m*-phenylenediamine. The products may be used as direct dyes for silk and wool and for cotton mordanted with tannin. They may also be used for coloring bacteria and for preserving foods from fermentation.

Azo dyes. I. G. FARBENIND. A-G. (Wilfried Hentrich and Rudolf Knoche, inventors). Ger. 522,201, May 15, 1928. The diazo comp'd from 4-nitraniline-2-sulfonic acid is coupled with 2-alkylaminonaphthalenesulfonic acids in which the alkyl residue contains more than 2 C atoms. An example is given. Cf C A 25, 2299.

Azo dyes. I. G. FARBENIND. A-G. (Leopold Laska and Arthur Zitscher, inventors). Ger. 522,293, July 26, 1928. Diazo tetrazo or diazoazo comp'ds not contg the $COOH$ or SO_3H group, are coupled with a 2,3-dihydroxy-naphthyl-4-amino-1-alkoxy-toluene. The dyes are insol in water, and are useful as pigments as well as for dyeing or printing vegetable fibers. Numerous examples are given. Cf C A 25, 2855.

Azo dyes. I. G. FARBENIND. A-G. FR. 698,219, June 30, 1930. Diazotized

aromatic amines are coupled with ω haloacylammonaphtholsulfonic acid and their derivs. Thus, a dye which dyes wool red shades is obtained by coupling diazobenzene with ω -chloroacetyl 1-amino-5-naphthol-3,6-disulfonic acid. Other examples are given.

Azo dyes. I G FARBENIND A G. Fr 697,540, June 17, 1930. Dialkyl ethers of aminohydroquinone or their derivs are diazotized and coupled with 1-benzoylamino-5-naphthol-4,6-disulfonic acid or its derivs substituted in the C₁₁ ring. Examples are given.

Azo dyes. I G FARBENIND A G. Fr 698,607, July 5, 1930. Azo dyes are prepd in substance or on the fiber by coupling diazo compds of esters of aminoterephthalic acid with aryl amides of 2-hydroxynaphthalene-3-carboxylic acid. Thus the dimethyl ester of aminoterephthalic acid is diazotized and coupled with the 2'-methoxy-1'-amide of 2-hydroxynaphthalene-3-carboxylic acid. Other examples are given.

Azo dyes. SOC ANON POUR L'IND CHIM & BAL. Fr 697,742, June 21, 1930. Aminodiarisulfones of the general formula R'R''N(H₂N)C₁₁H₁₀SO₂R, in which the amino groups are β to each other and R' represents H, alkyl, aralkyl or aryl, R'' is alkyl, aralkyl or aryl, and R is aryl, or ring substituents thereof, are treated with HNO₃ and coupled with arylides of β -hydroxynaphthoic acid in which the C₁₁H₁₀ nucleus may contain substituents such as halogen or OH and in which the aryl group may also be substituted.

Chromium compounds of azo dyes. I G FARBENIND A G. Fr 698,526, June 4, 1930. Azo dyes prepd by coupling diazotized aromatic amines containing either a OH or COOH group α to the amino group with 1-phenyl-3-methyl-5-pyrazolones containing in the phenyl group a OH and COOH α to one another are treated with Cr compds. Thus, the dye obtained by diazotizing α -aminobenzoic acid and coupling with (2-hydroxy-3'-carboxy-5-sulfo)-1-phenyl-3-methyl-5-pyrazolone is treated with a paste containing Cr₂O₃ and HCOOH. Other examples are given.

Vat dyes. FELICE BENSA. Fr 697,640, June 18, 1930. Perylenediketones dissolved in an appropriate solvent are electrolyzed. An example is given of the electrolytic treatment of dibenzoylperylene dissolved in concd H₂SO₄. The product dyes cotton in a blue shade which turns purple on exposure to the air.

Vat dyes. I G FARBENIND A G. Fr 697,875, June 25, 1930. Vat dyes of the N-dihydro-1,2,2',1'-anthraquinoneazine series are prepd by treating the dyes prepd by the process of Fr 693,919 (C A 25, 2004) with halogens or substances yielding halogens. Several examples are given.

Vat dyes. I G FARBENIND A G. Fr 698,345, July 3, 1930. The H₂SO₄ ester obtained by treating the leuco compd of 4,6-dichloro-6'-methoxybisthionaphthene-indigo (by condensing a 2 and of 2,3-diketohydro-6-methoxythionaphthene with 4,6-dichloro-3-hydroxythionaphthene) in the presence of a tertiary base with SO₃ or substances liberating it gives fast scarlet red shades on animal or vegetable fiber when used for dyeing or printing.

Vat dyes. I G FARBENIND A G. Fr 698,407, July 4, 1930. Two active derivs of 4-alkyl-5,7-dihalo-2,3-diketodihydrothionaphthenequinones are condensed with hydroxythionaphthenes with the exception of 4-methyl-5,7-dichloro-3-hydroxythionaphthenequinone. Thus, 4-methyl-5,7-dichloro-2,3-diketodihydrothionaphthenequinone-2 (β -dimethylamino)anil is condensed with 4-methyl-6-chloro-3-hydroxythionaphthene in glacial AcOH to give a product which dyes cotton in fast rose shades. Other examples are given.

Vat dyes. I G FARBENIND A G. Fr 698,638, July 8, 1930. Vat dyes of the N-dihydro-1,2,2',1'-anthraquinoneazine series are prepd by treating the products obtained by Fr 346,398, with halogens or agents liberating halogens. Examples are given.

Blue vat dyes. I G FARBENIND A G (Karl Thues, Theodor Meissner and Werner Zerweck, inventors). Ger 522,296, Aug 28, 1928. Reactive α compds of 4-methyl-5-halo-7-alkoxyisatins are condensed with 4-halo-1-naphthols. The reaction may be effected by warming the reagents together in PhCl soln. Examples are given. Cf C A 25, 2301.

Dye preparation. MARY B. PARRISH. Can 310,316, April 14, 1931. A dye compn consists of commercial aniline dyes, picrocratin, pancreas, acetic acid, citric acid, acetone, gelatin, tinctures of camomile flowers, hyoscy, ground cedar wood, and myrrh and dextrin.

Sulfuretted dyes. CHEMISCHE FABRIK VORM SANDOZ. Fr 698,583, June 30, 1930. See Brit 335,297 (C A 25, 1394).

Dibenzanthrone dyes. SCOTTISH DYES, LTD. Ger 522,406, Aug 27, 1927. See Brit 286,323 (C A 23, 283).

Xanthene dyes. IMPERIAL CHEMICAL INDUSTRIES, LTD. Ger 516,785, Mar 3, 1929. See Brit 314,825 (C A 24, 1518).

Acid dyes of the anthraquinone series. FABRIQUE DE PRODUITS CHIM. CI DEVANT SANDOZ 1r 699,306, July 2, 1930 Anthraquinone derivs having at least one replaceable group are condensed with halogenated aromatic amines, contg at least one halogen m- with respect to the amino group, and their substitution derivs and sulfonating the products thus obtained Thus, 1-amino-2-methyl-4-bromoanthraquinone is heated to 160° with *m*-chloroaniline and a small quantity of AcOK and crystd CuSO₄. The product is sulfonated Other examples are given Wool is dyed violet red, violet and violet blue shades.

Acid wool dyes. I. G. FARBENFABR. A.-G. (Walter Mieg and Kurt Bamberger, inventors) Ger 523,082, June 8, 1929 Amino- or alkylamino-1,1'-dianthraquinonyl amines are sulfonated with oleum in the presence of H₂SO₄ and Hg or its oxides or salts at 120-50° 1 examples are given Cf. C. A. 25, 2574

Alkylolamine salts of aliphatic acids and sulfonated fatty acids (dye assistants). JEAN G. KERN and CHARLES J. SALA (to F. I. du Pont de Nemours & Co.) U. S. 1,799,821, April 7, 1929 Dyeing is facilitated by mixing the dye with a product resulting from saponifying a higher fatty acid and glyceride with an alkylolamine such as mono-, di- or tri-ethanolamine, U. S. 1,799,822 relates particularly to the use of the product derived from oleic acid and an alkylolamine such as mixed ethanolamines U. S. 1,799,823 relates to the similar use of similarly formed derivs from stearic acid U. S. 1,799,824 relates to products formed from an alkylolamine such as ethanolamines and sulfonoleic acid or the product of sulfonation of castor oil U. S. 1,799,825 relates to the use in dyeing of these products U. S. 1,799,826 relates to printing pastes contg a lake of a basic color and an alkylolamine salt of a higher fatty acid

1-Amino-2-chloro-4-hydroxyanthraquinone. IVAN GUBELMANN, HENRY J. WETLAND and OTTO STALLMANN (to Newport Chemical Corp.) U. S. 1,799,156, March 31, 1929 This compd., a dye intermediate, forming dark red shining needles when recrystd from glacial HOAc, may be made by further heating a reaction mass resulting from the heating of 3'-amino-4',6'-dichloro-*o*-benzoylbenzoic acid with concd H₂SO₄ and contg 1-amino-2,4-dichloroanthraquinone to a temp. of 140-200° and maintaining the mass within that temp. range until evolution of HCl has practically ceased

Condensation products. I. G. FARBENFABR. A.-G. Fr. 608,525, June 3, 1930 Org. polycyclic N compds having at least 4 rings, particularly NO₂ or NH₂ derivs., are treated with org. compds having an open chain of at least 3 C atoms, in the presence of dehydrating agents, preferably in the presence of oxidizing substances favoring the reaction and finally substituting or condensing or both the reaction products thus obtained Thus, 6-aminobenzanthrone is condensed with glycerol in H₂SO₄. The product, 6(N),5-pyridinobenzanthrone, m. 250°, dyes acetate silk in green yellow shades, and by alkali fusion gives a product which dyes vegetable fiber a violet blue from the vat Many examples are given, including the prepn of pyridino-*allo* meso naphthodianthrone, m. 250-332°, a pyridinopyranthrone, m. above 300°, pyridino-3,4,8,9-dibenzopyrene 5,10 quinone, 4,3-pyridino-*Be*-dichloroanthraquinone 2,1-benzacridone and their derivs and condensation products The products have application in dyeing, printing and pharmacy

Aniline black dyeing. HIKOTO SATO Japan 89,540, Dec. 12, 1930 Fibers are treated first with a dil. soln of persulfon tannin, dried and then dyed with one of the following mixts (1) aniline salt 140, K₂Cr₂O₇ 150 and HCl 30 parts, (2) aniline salt 200, KClO₄ 40, CuSO₄ 80 parts, (3) white paste 1000, aniline salt 80, KClO₄ 40 and CuS 80 parts

Dyeing textiles. CAMILLE DREYFUS. Fr. 608,209, July 1, 1930 Textiles of or contg cellulose esters or ethers are dyed with vat dyes by applying the leuco compds by means of a medium of relatively low alk., e. g., 2-8 parts of alkali for 5 of dye Examples are given

Dyeing piece goods. AUGUST RÖMER, FÄRBEREI, BLEICHEREI UND APPRETURANSTALT (Alfred Linke and Walther Schramek, inventors) Ger 518,983, Feb. 10, 1927 A method of producing white or colored effects on piece goods during the sloppadded aniline black process is described Wool yarn is soaked with an ammoniacal or other alk. cascin soln to which oxidation retarding agents have been added, and is optionally after treated with CH₃O or acid Finally it is dried in moderate heat Thus, wool yarn is soaked in an aq. soln of cascin and NH₄OH and squeezed out It is then immersed in a bath of 1% aq. SO₂, dried and woven It remains unchanged in oxidation processes A further example is given

Dyeing acetate silk. SOC. ANON. POUR L'IND. CHIM. A. BALE. Fr. 607,541, June 17, 1930 Fast shades are obtained on cellulose esters and ethers by means of 1 alkyl-

amino-1 anilinoanthraquinones. Thus, the Me compd gives a light greenish blue shade on acetate silk.

Dyeing and waterproofing of cloth and fibers. YONETARO WASINO. Japan 89,546, Dec. 13, 1930. The dyeing and waterproofing are effected by one treatment with a mixt. of emulsified wax, paraffin, fatty acid, rubber or like substances in the presence of protective colloid, dyestuff and hydroxides of metals or their double salts.

Dyeing animal fibers. SOC. ANON. POUR L'IND. CHIM. A. BALE. Ger. 516,878, Dec. 28, 1929. The fastness of chromed azo dyes on animal fiber is improved by adding tannin and an aromatic hydroxycarboxylic acid to the dye bath, or by after-treatment with these substances. In the example, wool dyed with a chromed *o*-hydroxyazo dye is made fast to acid dressing and fulling by addn of tannin and salicylic acid.

Dyeing furs with "ursols." GOSUDARSTVENNAYA SUKREINO KRASILNAYA FABRIKA GOSUDARSTVENNAGO TRFSTA L'YNNAGADSKOI SHPINOT PROMISHLENNOSTI "LENINGRADORFZUDA." Russ. appl. 37,262, Dec. 19, 1928. Furs after being dyed with "ursols" by usual methods are first treated at ordinary temp with dil. solns. of $\text{Ca}(\text{ClO}_3)_2$, followed by a dil. sol. of H_2O_2 before their final finishing, whereby they are rendered harmless.

Coloring textiles. HENRY DAEVUS. Fr. 697,523, June 16, 1930. Textiles are colored by aq. suspensions or dispersions of free feuco compds of dyes. Several examples are given. Cf. C. A. 25, 2007.

Machine for dyeing cloth. MASCHINENFABRIK BENNINGER A.-G. Fr. 698,177, June 28, 1930.

Machine for dyeing cloth. MASCHINENFABRIK BENNINGER A.-G. Fr. 698,416, July 5, 1930.

Controlling means for dyeing machines. SIEMENS-SCHUCKERTWERKE A.-G. Fr. 698,252, July 1, 1930.

Beam-dyeing apparatus. WM. P. HORNBUCKLE and ROBERT F. CRAIG. U. S. 1,799,421, April 7, 1929. Structural features.

Apparatus for oxidizing mordanted cloth. TEINTURERIE CLÉMENT MAROT. Fr. 697,955, Sept. 30, 1929.

Textiles. HENRY DAEVUS. Fr. 698,418, July 5, 1930. Cotton or other cellulose fiber is eliminated from mixed fibers contg. animal fibers and (or) cellulose acetate or other esters or ethers of cellulose by heating the fibers in a soln. of a mineral acid or a salt having an acid reaction such as FeCl_3 and washing, etc.

Textile materials of organic cellulose derivatives. CAMILLE DAEVUS and HERBERT PLATT (to Celanese Corp. of Am.). U. S. 1,799,836, March 31, 1929. Org. cellulose derivs. such as cellulose acetate are prepd. contg. a substance such as a Sn compd. which acts as a weighting and swelling agent and gives increased resistance to heat.

Metallized fabrics. OSCAR TREBITSCH. Austrian 121,965, Nov. 15, 1930. Fabrics are treated with a metal salt soln. and then with a reagent capable of reducing the salt to the metal. After washing and drying, the metallic appearance is developed by mech. treatment with steel rolls or brushes in a lustering or beetling machine or the like. Cf. C. A. 25, 1392.

Treating cloth. HENRY DAEVUS. Fr. 698,181, June 28, 1930. Metallic radicals are incorporated in textile materials for mordanting, charging or dyeing by treating the materials with a current of particles of metal or oxide moving at a high speed, and oxidizing if necessary.

Vegetable fibers. HERFLEIN & CO., A.-G. Ger. 516,877, Nov. 29, 1927. In working up vegetable fibers by CuO-NH_3 soln. and NaOH , the fiber is treated first with CuO-NH_3 soln. and then directly with NaOH soln. The excess of the first soln. may be removed mechanically, e. g., by pressing, before the application of the second soln. The fiber is washed, dried, stretched, etc. The NaOH bath imparts a mercerized finish.

Threads for spinning. GIOVANNI MOSSO. Fr. 697,629, June 18, 1930. Fibers which may be spun are obtained by softening crude vegetable in water, treating chem. in an autoclave at above 130° in a bath contg. Na_2CO_3 1, NaOH 2-3 and olein or soap 0.5%. The fibers are afterward submitted to an energetic washing with water under pressure and afterward in a bath contg. soap and olein or Na sulfonate and dried.

Apparatus for sizing textile threads. JOSEPH ANNICO. Fr. 698,547, June 11, 1930.

Weavers' shuttles. I. G. FARBENIND A.-G. Fr. 697,496, June 16, 1930. Weavers' shuttles are made of wood impregnated with org. Cl compds. such as chlorinated $\text{C}_6\text{H}_5\text{Cl}$, or wax.

Scouring silk. I. G. FARBENIND A.-G. (Anne Nies-Harteneck and Roland Rilsch, inventors). Ger. 520,092, Mar. 6, 1928. Addn. to 513,373 (C. A. 25, 2578).

The method of Ger 513,373 is applied to undegummed crude silk in mixed fabrics or to hardened or colored but still undegummed silk. Examples are given.

Regeneration of natural silk. GAN'ITU KITA, GENTO YAMANOUTI, SYUZO MASUDA and INSTITUTE OF PHYSICAL AND CHEMICAL RESEARCH Japan 80,557, Dec. 15, 1930. $\text{Ca}(\text{NO}_3)_2$ soln of natural silk is coagulated by passing through 70% $\text{Na}_2\text{S}_2\text{O}_5$, 5H₂O soln. The regenerated fiber is further treated with AcONa soln or AcOH (1-2).

Preparing spinnable solutions from natural silk. I G FARBENIND A-G (Wilhelm Irion, inventor) Ger 522,046, Nov. 2, 1929. Degummed silk is dissolved in a concd strong mineral acid, e.g., 89% H_3PO_4 at a low temp., preferably below 0°, and the soln is treated with a salt soln, e.g., NaCl or HCOONa soln. The fibroin is thus pptd in a form readily sol at atm temp in solns of suitable alkali or alk earth metal salts, e.g., in 50% NaSCN soln, giving solns suitable for the manuf of fibers by spinning into baths of coagulating salts.

Preparing spinnable solutions from natural silk. I G FARBENIND A-G (Heinrich Fink and Ernst Rosner, inventors) Ger 522,408, Sept. 1, 1929. Addn to 510,489 (C. A. 25, 1103). In the method of Ger 510,489, about 0.02-0.1% of an ethereal oil is added to the acid spinning soln. The tendency of the spun fibers to stick together is thus reduced.

Twisting effect in cottonized fiber. A V TALANOV. Russ 10,408, Feb. 28, 1931. Fiber after having been cottonized by usual methods and after plucking and mellowing, is steamed in chambers with vapors of H_2SO_4 soln (8-10° Be) or HCl (12-15° Be) and then dried.

Preparing cotton goods for roughening. EDUARD TSCHÖRNER and HEINRICH RIEB, Ger 510,876, Sept. 10, 1929. The cotton fabric to be roughened is treated with an alk soap soln or with a sulfonated oil in alk soln to which an oxidizing agent has been added. An example of the latter agent is Turkey red oil and H_2O_2 .

Dulling cellulose esters and ethers. I G FARBENIND A-G (August Baumert and Richard Fischer, inventors) Ger 510,993, Nov. 12, 1929. Addn to 512,399 (C. A. 25, 1393). Cellulose ester or ether yarns or fabrics are uniformly dulled by impregnating with an aq soln of urea, and then steaming. The action of the urea soln may be enhanced by addn thereto of water-sol, aliphatic, hydroxy carboxylic acids, e.g., glycolic or lactic acid, or their salts. A dye may also be added to the urea soln. Examples are given.

Size-impregnated yarn package for use in knitting machines. JOSEPH A. HARROW, JAMES J. DIAMOND and GUSTAV BLEV. U. S. 1,799,230, April 7. Yarn packages are formed with convolutions of yarn of a type having an inherent tendency to snarl or loop in untensioned state, with the convolutions adhering to each other by the action of the size employed so as to resist being too readily unwound from the package.

Apparatus for the wet treatment of bank yarns. MASCHINENFABRIK TILLM GERBER SÖHNE & GERH. WANSLEBEN. Ger 522,382, Feb. 5, 1930.

Paste for textiles. HIKUZO KATO. Japan 90,201, Feb. 9, 1931. The paste is prepd by treating cherry-tree gum with NiOH , Na_2CO_3 , or K_2CO_3 for 24 hrs and then mixing with boric acid.

Apparatus for continuous steaming of material carried by looping bars. STANLEY H. FRANKLIN (to Textile Finishing Machinery Co.) U. S. 1,799,978, April 7. Structural features.

Washing liquid for fibers. HANS SCHWARZKOPF. Fr 698,101, June 26, 1930. A powd prepn suitable for making a washing or sprinkling liquid for treating wool or silk is made by mixing citric or tartaric acid with a non hygroscopic acid not acting on the citric or tartaric acid, such as boric acid or salicylic acid.

Waterproofing fibers. DUNLOP RUBBER CO., LTD. and THE ANODE RUBBER CO., LTD. Fr 697,730, June 20, 1930. Fibrous or cellulose materials are waterproofed by the application of rubber or vegetable resins in the form of an emulsion or aq dispersion with the addn of a wax or mixt of waxes.

Waterproofing cloth. EDMOND DECOAUX. Fr 698,691, Oct. 7, 1929. A coating of latex which may contain appropriate ingredients to give designs is used.

Treating textile fabrics with substances for waterproofing, mothproofing, etc. HILTON I. JONES. U. S. 1,799,047, March 31. Textile materials such as canvas are wetted with an emulsion contg petrolatum kept in suspension by the joint action of bentonite and an auxiliary emulsifying agent such as soap, casein and carrageen and the material is subsequently treated with a soln of a rare earth chloride such as Ce chloride so that a salt of the rare earth chloride is pptd in the material.

Apparatus for bleaching, washing or other wet treatments of fibers. EMIL GARY. U. S. 1,799,685, April 7. Various details are described of an app comprising

a treating drum of "frustoconical" form, having an inclined axis, with its lower side substantially horizontal, fibers and water, Cl soln., oil, acid, benzene or other treating liquid being delivered to the larger end of the drum and withdrawn from its smaller end

Apparatus for treating fiber packages under pressure with fluids such as bleaching, dyeing or washing liquids. OTTO J. OBERMAIER U S 1,708,623, March 31 Structural features

Drying apparatus for continuous strip material such as rubber-coated fabric. ROBERT R. JONES (to Firestone Tire and Rubber Co.) U S 1,709,375, April 7 Leakage of vapors at the inlet or outlet passageways is prevented by maintaining a body of inert gas about the passageway at a pressure greater than that within the apparatus and greater than atmospheric pressure Various details of apparatus are described

26—PAINTS, VARNISHES AND RESINS

A. H. SAMPSON

Irregular results of accelerated [paint] tests. P. NETTMANN *Farben-Chem.* 1, 103-4 (1930) cf C A 25, 2009—The composite nature of paint films is indicated, and an attempt is made to illustrate this by means of an equation Surface forces and thickness of coat, and the effects of temperature on these, are further influences acting against the possibility of obtaining uniform results from accelerated tests, even if such factors as vol., shape, wettability and degree of grinding of pigment could be standardized

B. C. A.

Protecting paints of aluminum. H. BUSCHLINGER *Hauszeit* 1° A II° u *Erstwerk A G Aluminium* 1, 237-43 (1929), *J. Inst. Metals* 43, 517—The properties are described of Al paints which consist of a mixt. of Al powder with a binding agent (linseed oil, varnish, cellulose lacquer, bituminous colors) Some examples of their application for rust prevention are given

G. G.

Protective coatings of aluminum paint. H. BUSCHLINGER *Vorträge gek. Jahresversammlung 1929 zu Wien* (Suppl. to *Korrosion u. Metallschutz*) 1929, 49-53, *J. Inst. Metals* 43, 517—The value of Al paint as a protection against corrosion is discussed, it is particularly resistant to the action of moisture and SO₂ and, if desired, may be colored by the addn. of pigments The manifold properties and uses of the paint are described, it reflects 70% of the incident light and heat waves, and its radiation power is only 30% of that of a perfect black body One kg. of the paint will cover about 15-20 sq. m. of surface satisfactorily

G. G.

Chalking of paints. IV. Practical studies of the chalking of oil paints. RICHARD KEMPE *Farben-Ztg.* 36, 1173-6, 1216-7, 1256-9 (1931), cf C A 25, 1104—The chalking of paints made up with one or more of the pigments, (1) red seal lithopone, (2) bronze seal lithopone, (3) low oil-absorption lithopone, (4) carbonate white lead, (5) sublimed white lead, (6) Kronos A titanium pigment, (7) TiO₂, (8) ZnO, (9) lead chromate, in vehicles contg. one or more of the oils, tung, boiled linseed, linseed and tung standoils, was studied The area under the time-degree of chalking curve was taken as the amt. of chalking The type of vehicle proved to be unimportant The chalking of the single pigment paints was 4 least, 1, 7 and 6 increasing in the order named The chalking of paints contg. lithopone in combination with 9, 8, 4 or 5 was 9 least, followed by 8, 4 and 5 Cracking sometimes accompanied pronounced chalking, e. g., lithopone paints, but slight chalking was not always accompanied by cracking, e. g., white lead paints

G. G. SWARD

Analysis of an oil color. A. GARCIA BANCOS. *Anales soc. espal. fis. quim.* 29, 182-3 (1931)—A violet oil paste defied sepn. because the dyestuff dissolved in all known org. solvents and was insol. in water Sapon of the oil gave hopeless emulsions. Since such dyestuffs have usually a base of triphenylmethane, HCl was passed in, pptg. the dyestuff completely The oil was dolphin or whale oil

E. M. SYDNES

Human factors in color judgment. W. O'D. PIERCE *J. Oil & Colour Chem. Assoc.* 14, 90-109 (1931)—A review of the psycho-phys. aspect of color matching indicates that individuals vary widely in the 3 fundamental attributes of color Normality in one respect does not guarantee normality in all others Results on pigments do not necessarily parallel those on lights There are probably as many supernormals as subnormals It is further emphasized that color judgment depends upon the human factor and that color has no existence apart from the presence of the observer It is shown that experience is a large factor in matching color

G. G. SWARD

A simple device for measuring gloss. G. G. SWARD AND S. A. LEVY. *Am. Paint*

Et Varnish Mfrs'. Assoc., Circ. No 380, 162-4(1931)—A low cost device for viewing the reflection of an illuminated slit 1 X 20 mm at an angle of incidence and reflection of 45 is described and illustrated. Illumination is by an ordinary flashlight. The reflection from a sample is merely compared with that from standards of known gloss. The range of gloss from that of polished plate glass ($\approx 100\%$) to 0% is divided into 5 steps: above 90, excellent, from 80 to 90, good, from 25 to 80, fair, from 10 to 25, fair, below 10, none.

G. G. SWARD

Gloss and its quantitative measurement. E. D. RIES AND C. B. GILBERT. *Proc. Am. Soc. Testing Materials* 30, Pt. II, 911-23(1930)—An app. is described for measuring gloss by photometric detn. of the intensity of specularly reflected light. The reflected light goes through a sharp max. when the angle of incidence of the rotating panel board equals the angle of reflection. Flat surfaces show no max. The amt. of specularly reflected light increases with the angle of reflection and approaches 60% for some polished panels. Gloss changes markedly with the angle of incidence except in very glossy panels. Diffuse reflection from darkly pigmented lacquered surfaces has been shown to be of the same order of magnitude as that from light ones. H. M. S.

Application of microchemical methods to investigation of the pigments in paintings. H. HERRMANN. *Mikrochem. Emich. Festschr.* 1930, 182-62—Particles of pigments may easily be removed from oil paintings by means of a fine needle moistened with glycerol, without causing any visible injury. Details are given of microanalytical methods suitable for the examn. of the pigments. B. C. A.

Report of cooperative work on the determination of hiding power of paints and pigments and the tinting strength of pigments. ANOV. *Proc. Am. Soc. Testing Materials* 30, Pt. I, 748-52(1930)—A series of pigment samples and pigment-in-oil mixts. was sent out by Am. Soc. Testing Materials, subcommittee VIII on Methods of Analysis of Paint Materials, to cooperating institutions who were requested to make tinting strength and cryptometer tests. The results led to the conclusions: (1) that the former Tentative Method D282-28T for tinting strength did not give consistent results, (2) that tinting strength should be studied further by the committee; and (3) that hiding power can be measured satisfactorily by the cryptometer, reproducible results being obtained by the several laboratories reporting on the pigment-oil mixtures.

H. M. STARK

A discussion of hiding power and its measurement. G. P. A. STUTZ, JR., AND G. S. HASLAM. *Proc. Am. Soc. Testing Materials* 30, Pt. II, 884-90(1930)—Indirect measurement of hiding power by measuring related properties such as tinting strength is not satisfactory. The paint-out test for hiding power is correct in principle but is time-consuming in application. Improvements necessary in the test before it will meet the definition of Subcommittee VIII on Methods of Analysis of Paint Materials include more uniform application, possibility of increasing the film thickness by an amount less than a normal coat, etc. The Pfund cryptometer measures hiding power simply and accurately but is applicable only to wet paints and to paints of over 70% brightness. The Pfund black-and-white cryptometer gives an accurate measure of the wet hiding power of a paint of any brightness. The Haslam spinning-film method measures hiding power accurately but requires expensive app. The modified paint-out test of Gardner is less accurate but simple and inexpensive. H. M. STARK

Hiding power and tinting strength of white pigments. R. L. HALLETT. *Proc. Am. Soc. Testing Materials* 30, Pt. II, 895-910(1930)—This paper describes a method for detg. hiding power and tinting strength for white pigments. The hiding power test consists of a brush out of several coats on a white board with a black stripe down the center. Tinting strength was detd. by detn. of the amount of carbon black necessary to dil. the white color of the unknown to that of a standard white contg. a fixed amount of black. It is important in measurement of tinting strength that the pastes all have the same consistency and the same brightness. Hiding power and tinting-strength values for 20 white pigments obtained by the methods described showed a straight line relationship to each other. The hiding power in sq. ft. per lb. was equal to its tinting strength (carbonate white lead $\approx 100\%$) multiplied by $0.10 + 3$.

H. M. STARK

Manufacture of titanium white. HANS J. BRAUN. *Metallbörse* 21, 507-8(1931)—Ilmenite is treated with 90% H_2SO_4 at about 80° or with 80% acid at $180-250^\circ$. $Ti(SO_4)_2$ and $Fe_2(SO_4)_3$ are leached from the brown mass resulting from the attack. Ti is pptd. as $Ti(OH)_4$ by hydrolysis of the soln. Hydrolysis is influenced by concn., temp., H^+ ion concn., manner of agitation and the presence of foreign substances, both org. and inorg. The product of hydrolysis is amorphous and contains some SO_4 , not removed by washing. The H_2SO_4 is neutralized by $BaCO_3$ or $CaCO_3$. Calcination

produces TiO_2 of microcrystalline form. Other white pigments are mixed with TiO_2 to give the product special properties. TiO_2 pigments are superior in tinting strength and covering power and are chemically inert.

H. M. STARK

The inexhaustible red lead problem. HANS HENNINGSEN. *Farben-Ztg.* 36, 1127-30(1931)—The value of red lead as a pigment is undoubtedly due to its ability to form so-called metallic glyceride systems, a property not possessed by iron oxides. Basicity or ability to form metallic soaps is not related to the durability of films. G. G. S.

Reaction of litharge with linseed oil at room temperature. K. CHARISIUS AND E. KINDSCHER. *Farben-Ztg.* 36, 1297-1300(1931)—A mixt. of raw linseed oil 500 g. and litharge 750 g. was placed in a completely filled, closed flask and shaken daily for 7 weeks. At the end of that time there was obtained, (1) 316 g. of a clear acetone sol. oil contg. 2.6% Pb. The residue yielded an ether ext. (2) an org. compd. contg. 46% Pb. The residue from 2 was extd. with a mixt. 9:1 of benzene and alc. and yielded 47 g. of a solid dark brown compd. contg. 50.4% Pb. The residue from 3 contained no org. material. The principal change accompanying the reaction was a slight increase in the acid no. of the clear oil (1) from 5.6 to 7.1, and the oxidation (by the litharge) of the org. portion of (2). The compn. of the lead compd. in (1) is unknown but upon standing a white compd., apparently a neutral Pb soap of satd. fatty acids, is deposited.

G. G. SWARD

Lithopone. MASAO UCHINO AND K. KAMITANI. *Repts Osaka Imp. Ind. Research Lab. Japan* 11, No. 3(1931)—A general review is given on the methods of prepn., properties and the methods of analysis.

F. I. NAKAMURA

The preparation of ultramarine blue and health. G. ISSOGLIO. *Industria chimica* 6, 27-30(1931)—The manuf. of ultramarine blue, because it uses as raw materials silica, kaolin and S, developed a great deal of dust, which is found to induce pulmonary diseases. This industry is, therefore, classed as hazardous, and it is recommended that the dust be kept under control.

A. W. CONTIERI

The preparation of an antimony yellow. JAROSLAV MILBAUER. *Chem.-Ztg.* 55, 222(1931)—A bright yellow product is obtained by grinding together Sb_2S_3 and 25% NaOH (1:6) in a ball mill for 3 hrs. The product is a cold soln. of Sb_2S_3 and Sb oxides. This pigment is stable to 200°, has good covering power, but may not be mixed with colors that react with sulides.

W. GORDON ROSE

Manufacture and evaluation of Schweinfurtergreen. A. TRITAMER. *Farben-Chem.* 1, 114-6(1930)—The prepn. and properties of Schweinfurtergreen (Cu arsenoacetate) are described. The characteristics of this pigment depend on purity and regularity of particle size, and details of its evaluation by chem. analysis and by sieving and settling tests are given.

B. C. A.

Precipitating agents used in color-lake manufacture. A. W. C. HARRISON. *Farben-Chem.* 1, 105-8(1930)—An account is given of insol. agents and mordants functioning by occlusion, e. g., green earth, artificial Fe silicate, Al_2O_3 , of sol. metal salts (true precipitants), e. g., $\text{Ca}(\text{OAc})_2$, $\text{Pb}(\text{NO}_3)_2$, BaCl_2 , and of special synthetic agents, e. g., Tamol NN and Katanol.

G. G.

Wood stains. A. C. HEINEMANN. *Farben-Ztg.* 36, 1259-60(1931)—Acid dyes have but little affinity for cellulose and when used for dyeing wood their effectiveness is due to the presence of lignin. This fact was demonstrated by being able to dye newsprint paper slightly but filter paper not at all with Orange G. Basic dyes are more suitable for wood. The staining of wood with particular emphasis on wax emulsion stains is briefly discussed.

G. G. SWARD

Requirements of pigments for the linoleum industry. H. G. BODENBINDER. *Farben-Chem.* 2, 58-61(1930)—The desirable qualities of linoleum pigments are fastness to light, resistance to cleaning agents, heat and acid; freedom from oxidation catalysts and from sol. Fe salts. Each of these is discussed, and a general description of linoleum manuf. is given.

B. C. A.

Notes on the oxidation of metallic keels and the means of protecting them from rust. H. MASSEILLE. *Peintures pigments vernis* 7, 1372-5(1930), 8, 1401-4, 1430-3(1931).—Devices for grinding and mixing paints are discussed.

J. W. PERRY

Manganese soaps. H. HANS WAGNER AND G. HOFFMANN. *Farben-Ztg.* 36, 1214-6(1931), cf. C. A. 25, 1695—The degree of swelling of linseed oil films in water is proportional to the amt. of Mn soaps present. Films contg. Mn hydroxide, umbers, Mars yellow and Mars yellow plus 5 and 10% Mn soaps were prepd. Only the film contg. pure Mars yellow was satisfactory from the standpoint of swelling in water.

G. G. SWARD

Cooling produced by evaporation. HERBERT DABISCH. *Farben-Ztg.* 36, 1300(1931)—A Lambrecht psychrometer was used to obtain the lowering of temp. pro-

duced by the evapn. of some 21 solvents used in the paint industry. The wet balls was unventilated and the detns. were made at 27° and 10°, with relative humidities of 63% and 91%, resp. With hygroscopic liquids, instead of a lowering, a rise in temp. took place. This was due to the heat of soln. exceeding the heat of evapn. G. G. S.

The formation of films of drying oils. A problem in organic colloidal chemistry. C. P. A. KAPPELMFIER. *Chem. Werkblad* 28, 174-83 (1931).—A lecture, reviewing the various theories which have attempted to explain the mechanism of the drying of oils. Stress is laid upon the fact that the problem must be considered not simply as one in org. chemistry, but rather as one in org. colloid chemistry. G. CALINGAERT.

Ursolic acid and other plant products (in lacquer and varnish). H. A. GARDNER, G. G. SWARD AND A. W. VAN BRICKENRATH. *Am. Paint & Varnish Mfrs' Assoc. Circ.* No. 379, 154-60 (1931).—The drying times of varnish or nitrocellulose lacquers in which ursolic or oleanolic acids have been added are greatly extended. The acids dissolved in warm varnish are pptd. when the varnish is cooled and act as flattening agents. The tensile strength and elongation of lacquers are slightly reduced but the gloss is greatly increased by the presence of the acids. The possibilities of derivs., such as esters and salts as plasticizers are suggested. G. G. SWARD.

Mirror-surface lacquers. MAX URMES. *Diamant* 51, 286-8 (1923). *Ceram. Abstracts* 9, 268.—The properties of some of the varnishes or lacquers used in the mirror industry are discussed. Varnishes containing silver are not used for silver mirrors as the residual silver causes them to be hygroscopic to some extent, while benzene varnishes are not good in cold weather, since the resin is decomposed and drying is held up. No alternative solvents are suggested. G. G.

Thinners for cellulose lacquers. P. S. SYMONS. *Brit. Plastics* 2, No. 23, *Synthetic and Applied Finishes* 2, 57, 12 (1931).—The methodical blending of thinners for nitrocellulose lacquers are outlined with particular respect to such points as "flow," blushing, pinholing and wrinkling of the lacquer films. Tables are given showing the diln. ratio for nitrocellulose in various solvents using benzene, xylene, solvent naphtha and kerosene, resp. as diluents and the evaporative rates for a no. of solvents, also formulas for a general high flash point thinner, a general low flash point thinner and a cheap thinner. Sy warns against the use of "universal thinners." D. THULESEN.

Butyl and amyl lactates as solvents for cellulose lacquers. OTTO GERHARDT. *Chem. Ztg.* 55, 222 (1931).—A brief discussion of the properties that characterize butyl and amyl lactates as desirable cellulose solvents. W. GORDON ROSE.

Plasticizers for nitrocellulose lacquers. P. S. SYMONS. *Brit. Plastics* 2, No. 22, *Synthetic and Applied Finishes* 1, 119-71 (1931).—The proper choice of plasticizers is discussed with consideration of the various purposes for which the lacquers are to be used. The main characteristics of the following principal plasticizers are described: diamyl phthalate, dibutyl phthalate, methylcyclohexanol oxalate, terecyl phosphate, triphenyl phosphate, Sipalin ADAS (methylhexalin ester of adipic acid), Sipalin MOA (methylhexalin ester of methyladipic acid), camphor, flu. stearate, castor oil, linseed oil, tung oil and rapeseed oil. D. THULESEN.

New resins for the cellulose-lacquer industry. F. HARRISS. *COTTON Brit. Plastics* 2, No. 21, No. 22, *Synthetic and Applied Finishes* 1, 157-8, 163-8 (1931), cf. *C. A.* 24, 125.—The manuf. of nitrocellulose lacquers and enamels is broadly described. Varied improvements in the properties of these have been obtained with the use of gutta percha and balata resins. D. THULESEN.

Resin studies. I. Preparation and autoxidation of precipitated lead rosinate. W. A. LA LANDE, JR. *J. Am. Chem. Soc.* 53, 1858-68 (1931).—An app. is described in which all the steps of the prepn. of Pb rosinate (I) may be carried out in an inert atmo. 2% solns. of Pb(OAc)₂ and Na rosinate were found most practical for the pptn., the compn. of the ppt. was also found to vary with the rate of addn. of the Pb(OAc)₂ to the Na rosinate and with the time and rate of agitation of the reaction mixt. Details are given of the prepn. and a method is proposed for the analysis of the metal content of rosinate. The rate and degree of O absorption by I depend, for a given pressure, on the temp. The amt. of O absorbed at equal corresponds to 3 atoms of O, the reaction product at this point is probably a combination of reaction products, rather than a definite compd. of 3 atoms of O with 1 mol. of Pb abetate. The mixt. slowly liberates I from aq. KI quickly from acidified iodide. The product obtained at 60 atm. pressure also liberates I from aq. and acidified iodide. The dissocn. of the end (or equiv.) product in the autoxidation of I evidently proceeds progressively as the temp. is raised, there is an indication that this dissocn. is never complete, and that a max. dissocn. is reached when 1 at. equiv. of O remains fixed by the rosinate. The amt. of O absorbed in each case is thus an index of the stability (and the comen.) of the end product at a

given temp for a const pressure. The results are discussed from the viewpoint of the theory suggested by Dupont and Lévy (*C A* 24, 875, 1786, 1621, 2315, 5214) for the autoxidation of aliphatic acid. C. J. WEST

Resin studies. VIII. The Donath resin test. FATCH STROCK. *Farben-Ztg* 36, 1176-7 (1931), cf *C A* 25, 222—The Donath test (*C A* 24, 5515) was carried out on a no. of natural and synthetic resins and varnishes with and without rosin. One g. of powder (in the case of resins) was cooked with 5 cc HNO₃ (sp gr 1.323), cooled, diluted with an equal amt. of water and finally neutralized with a large excess of NH₃. In many cases the test is useful, but it does not definitely identify rosin. G. G. S.

Gum resin production and dry distillation. I. BONAHOV. *Trudovi Vsesoiuznogo Nauchno-Issledovatel'skogo Instituta Khimicheskoi Promyshlennosti* 3, 74-80 (1929)—A number of processes resulting in change of products take place in gum resin production. The construction of apparatus permitting the primary products to escape before further decomposition is of great importance. B. does not agree with the theory of Clason that the primary products are volatile in high vacuum. The retort used for gum resin production in the Viatka region has always an exit for non-volatile but fusible products before they are decomposed by heat. These are not obtainable in ordinary dry distn. of wood and B. therefore differentiates the latter from thermal decomposition of wood, which is the process used by him. A number of expts. with cellulose, lignin, sawdust, combustible slate, and peat were carried out in which the material was heated to 215-75°. Lignin gave mostly products non-volatile in a high vacuum; half of them are sol in Et₂O, the rest insol in the usual org. solvents; sol in NaOH; less sol in soda. As to the ether sol products it is known only that they are a mixt. of weak acids and phenols. Cellulose gave water sol products (except charcoal) of sweet-caustic taste, 42-55% of the cellulose products were non-volatile. They reduced an ammoniacal soln. of silver oxide, gave a cuprous ppt. with FeHCl soln. and showed reaction for carbohydrates with alkalis and with phenylhydrazine. First they could not be fermented but after treatment with dil. acids they did. The volatile matter contained considerable quantities of furfural and hydroxy-methylfurfural. Wood and peat gave similar products. The process is not changed by application of steam and vacuum but the yields are increased. Lign. would probably give also solid infusible products, but for technical reasons this was not tried. J. G. TOLPIN

The artificial synthetic resins. RPNF. CURSANTAU. *Bull. soc. ind. Rouen* 58, 302 (1930)—A review is given of the methods of production, properties, and uses of synthetic resins, classified as phenol-formaldehyde, urea-formaldehyde, glycerophthalic (glyptal) and coumarone. I. H. ORILL

Phenol- and cresol-formaldehyde resins. HANS STÄGER. *Helv. Chim. Acta* 14, 285-304 (1931)—Resins were prepd. by condensation of HCHO and PhOH (I) or MeC₆H₄OH (II) and mixts. of the 2. The II used contained varying quantities of p and m isomers. The properties of these resins were compared. The sp. cond. and loss in dielectric rise while the dielectric const. falls with increase in I content. To a smaller extent, the reverse is true for the II resins with rise in the m compd. These facts show that polymerization takes place to a higher degree with II. A fall in the dielectric const., etc., is obtained by heating the resins. The volatilities in a/c and NaOH also drop with increase in duration of heating. The force of cohesion (cementing force) drops to a min. upon heating and then rises steadily as the heating is continued, with increase in I the min. is lower and a longer time is required to attain it. I. M. L.

The electrolytic preparation of white lead (SACHSE) 4. Some unusual properties of colloidal dispersions (WILLIAMSON) 2. Hydroxide systems in Fe oxide colors (WAGNER) 2. New compounds of Ti (GARDNER, BIFLOUSS) 10. Polymerizing oils and fats (for oleum industry and for addition to varnishes) (Fr. pat. 697,785) 27. Ornamenting artificial oil cloth, etc. (Austrian pat. 121,999) 29. Condensation products [resins] (Fr. pat. 697,700) 22. Azo dyes [for use as pigments] (Ger. pat. 522,205) 25. Plastic materials (Fr. pat. 697,471) 23.

DR. WILD, A. MARTIN. *Naturwissenschaftliche Gemäldeuntersuchung*. Munich: B. Heller. 101 pp.

Paint mill. ELMER PETERS (to Kent Machine Works, Inc.) U. S. 1,709,627, March 31.

Titanium white. ALEXANDER FOLLIOT and NICOLAS SAINDRECHIN. Fr. 698,510, May 26, 1930. Titanium white is prepd. by sepn. of the metal from its ores, e. g.,

aliments by concomitant reduction and oxidizing volatilization under the action of an air blast heated from 750° to 850° acting on a thin layer of a mixt. of powd. ore and reducing fuel moving forward under the action of air jets.

Zinc oxide. MANUFACTURES DE PRODUITS CHIMIQUES DE JOUY-EN-Josas (Anciens Établissements Louis Descamps) Fr 697,760, June 21, 1930 ZnO free from impurities is made by transforming the Zn starting materials into Zn(AcO)₂ either by soln. in AcOH or by double decompn., eliminating the impurities and submitting the solns. or crystals of Zn(AcO)₂ to pressure and temp. to allow the disson. of acetate with recovery of AcOH.

Destructively distilled castor-oil condensation product suitable for use in paints, varnishes and lacquers. EDWARD C. HOLTON (to Sherwin Williams Co.) U. S. 1,799,420, April 7. Castor oil is subjected to destructive distn. until the residue is 60-80% the wt. of the original oil, and the residue is heated with 0.5-3.0 times its wt. of glycerol until the esterifiable constituents are substantially esterified, and the product thus obtained is then heated with 0.75-5.0 times the wt. of the residue, of phthalic anhydride until condensation is effected and a clear homogeneous product is obtained.

Printing inks. JOHANNES SCHREIBER Ger. 622,277, Nov. 6, 1928. For the prepn. of odorless printing inks, a vehicle is used consisting of, or comprising, the viscous residues from the vacuum distn. of castor-oil acids.

Printing inks. JOHANNES SCHREIBER Ger. 622,480, Nov. 6, 1928. For the prepn. of odorless printing inks, a vehicle is used consisting of, or containing, an ester of a polyhydric alc. with 9,11-octadecadiene-1-carboxylic acid. This acid is prepd. by the vacuum distn. of castor-oil acids. The esters may be boiled to a suitable consistency before use. Cf. preceding abstr.

Printing ink. I. G. FARMER AND A. C. Fr. 698,639, July 8, 1930. Varnishes, which have a basis of cellulose derivs. obtained by using only hydrophobic constituents, are used in flat printing machines. Examples are given of the use of a varnish made by mixing nitrocellulose of low viscosity with castor oil, phthalic acid ester, benzoic acid ester, dammar resin or an esterified resin and lamp black or ZnO.

Printing ink. GOOZO NAKAO, Japan. 60,225, Feb. 10, 1931. The ink contains KI or NH₄I and starch.

Coating composition. EDWARD F. ARNOLD (to Canadian Industries, Ltd.), Can. 310,845, Apr. 28, 1931. A striping enamel comprises nitrocellulose of low viscosity, a pigment and a polyhydric alcohol polybasic acid resin modified by more than 80% of drying-oil component. Cf. C. A. 25, 834.

Nitrocellulose solvent and coating composition. DAVID CARNEGIE, JR. (to The Atlas Powder Co.) Can. 309,861, Mar. 31, 1931. A coating compn. comprises 20 oz. of cellulose nitrate per gallon of soln. The solvent comprises by vol. 30 parts of AcOEt, 35 parts of EtOH and 35 parts of a non-toxic hydrocarbon oil fraction corresponding to a naphtha having an initial b. p. of about 20°, with 50% passing over at about 90° to 95°, and 90% at about 105° with a dry point at about 115°.

Synthetic-resin coating composition. MORRIS H. HOPKINS (to The Canadian Industries, Ltd.) Can. 310,124, Apr. 7, 1931. A coating compn. for a linoleum base contains as ester of glycerol, a polybasic acid, acids formed by the hydrolysis of linseed oil and acids formed by the hydrolysis of China wood oil, and a pigment.

Pyroxylin solutions. HAMMILL PAPER CO. Fr. 697,780, June 23, 1930. See U. S. 1,768,253 (C. A. 24, 4391).

Varnish. "CIXINE-WEKKE" JOSEF LORENZ & Co. Austrian 121,984, Nov. 15, 1930. Varnishes comprising wax, fatty oils, a Co. siccativ and beozine or turpentine are prepd. by emulsifying a part of the wax, after preliminary comminution, with the fatty oil in the cold, and adding to the emulsion a cooled soln. of the siccativ in the fatty oil and a soln. of the rest of the wax in benzine or turpentine. The wax may be paraffin, ceresin or beeswax, and the fatty oil may be wood oil, linseed oil or thickened linseed oil (stand oil). An example is given. Cf. C. A. 25, 2311.

Varnishes. BAKELITE CORP. Fr. 698,091, June 26, 1930. A resin soln. contains a non-hardening resin having a basis of thiourea, an anhyd. hardening agent contg. CH₂ in amt. sufficient to harden the resin and a solvent for the resin and hardening agent. The soln. is used as a varnish. Cf. C. A. 25, 2866.

Varnishes. MAURICE PRACOT, Fr. 698,712, Oct. 9, 1929. AcOH is added to cellulosic varnishes to increase the stability of the colors obtained by incorporating artificial dyes therewith.

Crystallizing varnish. WILLIAM O. STAUFFER (to Canadian Industries, Ltd.) Can. 310,846, Apr. 28, 1931. A crystg. varnish contains 19 parts by wt. of a resin made from glycerol phthalic anhydride and China wood-oil acids, 33 parts of blown

China wood oil, 5.5 parts of Co drier soln., 9.5 parts of solvent naphtha and 28 parts of high boiling gasoline

Plant for making varnish from wood oil. WILHELM SCHMIDING Ger. 522,407, Feb. 14, 1925.

Recovery of values from lumsa in varnish making. ROBERT S. PERRY. U. S. 1,799,177, April 7. The fumes are passed through a treatment conduit having an open end, and an alk. liquid is sprayed within the conduit and toward its open end, condensates formed are collected with the alk. liquid, and after sepn the alk. liquid is further used for spraying. App. is described.

Lacquers, etc. I. G. FAHREND A.-G. Fr. 697,348, June 13, 1930. Lacquers and plastic materials having a nitrocellulose basis are colored an intense greenish blue, fast to light by the dyes obtained by condensing the hydrochloride of a dialkyl nitro-*o*-aniline with a dialkyl *m*-aminocresol or a monoalkyl *m*-aminophenol. Thus, 2 di-methylamino-2'-diethylamino-3'-methyloxazine may be used. Other examples are given. Cl. C. A. 25, 2860.

Pigmented lacquer base. HENRY J. HEMINGWAY and WM. A. WEIDLICHER. U. S. 1,798,840, March 31. A non aq. mixt. of nitrocellulose and unground pigment such as "titanox" is subjected (without grinding or rolling pressure) solely to the kneading action of a mixing machine in the presence of a relatively small quantity of a liquid volatile solvent such as alc., LiOAc and CaI₂ to imitate the dissolving of the nitrocellulose simultaneously with the pulverization of the pigment and its dispersion throughout the batch (the amount of the solvent being insufficient to dissolve all the nitrocellulose but sufficient to produce a high viscosity during the kneading), and addnl solvent is added after the pigment is broken down and dispersed.

Fire-resisting lacquer. FERNANDO S. VIVAS (to International Fireproof Products Corp.) U. S. 1,800,120, April 7. Gum dammar, "soluble cotton," oil such as castor oil, CaCl₂, H₂BO₃, alc., BuOAc, EtOAc, toluene and perchloromethane are used together.

Condensation products of phenols and aldehydes. WILFRED S. ROTHERA, STANLEY BLYTHEN and H. R. GILLESPIE (Alphous Osterseizer and Franz Riesenfeld, inventors) Ger. 516,677, Mar. 13, 1928. To produce the condensation product as a powdery ppt., a basic condensing agent contg. less than 10% of a dispersing agent or protecting colloid, and which causes a 25% increase of volume, is added. Neutral salts may be present. Thus, a mixt. of *m*-cresol and CH₃O is added to distd. water and gum arabic and heated. A yellowish sand like resin powder results. Further examples are given. Cl. C. A. 23, 1180.

Phenol-aldehyde condensation products. BAKELITE G. M. B. H. (Fritz Seebach, inventor) Ger. 516,760, July 18, 1928. Pure resin like PhOH-CH₂O condensation products are prepd. by dissolving the crude condensation products in aq. solns. of readily sol. salts of PhNH₂ or toluidine with acids such as HCl, HBr, HNO₃ or CH₃COOH, and distg. the soln. The distn. may take place *in vacuo* or with addn. of superheated steam. In an example, aniline-HCl water and cresol-CH₂O condensation products are distd. Further examples are given. Cl. C. A. 24, 3662.

Moldable mixtures. BAKELITE CORP. Fr. 698,183, June 28, 1930. A moldable mixt. contains a resin of the hexamethylenetetramine, triphenol type, an org. retarder such as colophony and a filler such as mica or a hardened and divided resinoid. Cl. C. A. 25, 2823.

Resinic acid derivatives. CHEMISCHE FABRIKEN KURT ALBERT G. M. B. H. Fr. 697,470, June 16, 1930. Derivs. of natural or artificial resinic acids are prepd. by heating resinic acids with acid derivs. b. above 200°, the acid of which when liberated may be easily eliminated by phys. or chem. means. Examples are given of the heating of colophony and of various copals with dibutyl phthalate, benzyl acetate, triglycero-stearin or acetamide.

Resins. JOH. M. WEISS. Fr. 698,623, July 7, 1930. See Brit. 338,815 (C. A. 25, 2581).

Fractionating rosin into resene and abietic acid. FREDRICK W. KRESSMAN. Can. 310,774, Apr. 28, 1931. Rosin is distd. at a pressure not greater than 10 mm. of Hg. Between 205° and 240° the vapors contain principally resene. Between 240° and 260° abietic acid chiefly distils over.

Synthetic resins. GEORG WALTER. Austrian 121,999, Nov. 15, 1930. Resinous products are prepd. by warming with bases monomethylolurea or monomethylolthiourea, or reagents yielding them, or the products obtained from them by splitting off water, in the absence of any substantial amt. of solvent. Thus, monomethylolurea may be mixed with 1-7% of NaOH, added either as a solid or in concd. soln., and the mixt. heated to 100-130°. The base may be neutralized at the end of the reaction. The

products can be hardened by heat, with or without pressure, and may be used as molding powders or liquid ingredients. Examples are given.

Synthetic resins. I. G. LARSEN AND A. G. (Leo Rosenthal, inventor) Ger 520,834, Apr 18, 1929. Adds to 517,477 (C A 25, 2317). Compds of BF_3 with aliphatic hydroxy acids, or with unsat'd aliphatic acids, are used in the process of Ger 517,477.

Synthetic resins. BAKFLITE CORP. Fr 697,525, June 16, 1930. A moldable material is made by treating fusible resinous products from the condensation of PhOH , with substances, such as $\text{Ca}(\text{OH})_2$, which transform the PhOH into phenolate. The cooled or solidified product is mixed with a comp'd contg active CH_2 groups to transform the fusible resin into an invol infusible resin. Appropriate filling materials may be added. Cf C A 25, 2867.

Synthetic resins. BAKFLITE CORP. Fr 697,753, June 21, 1930. The hardening of a phenol furfural resin is accelerated by the addn. of a basic substance such as lime and a comp'd contg CH_2 , such as $(\text{CH}_2)_4\text{N}$.

Synthetic resins. LUCITY SELLER. Fr 697,968 Oct 1, 1929. The OH groups of compds such as glycerol are sated with natural resins of acid function (colophony, copal, etc.) and a dibasic acid (adipic). Thus, glycerol 85, colophony 150 and adipic acid 140 g. are heated together.

Synthetic resins. ILLIS-FOSTER CO. Fr 698,598, July 3, 1930. Resins which may be dissolved or dispersed in water are made by treating the condensation products of a polyhydric alc. with a polybasic acid, with or without monobasic acids, by a base such as an oxide or hydrate of Ca, Ba, Mg or Zn or NiH , or an org. base such as $(\text{CH}_3)_3\text{N}$. Cf C A 25, 2312.

27—FATS, FATTY OILS, WAXES AND SOAPS

E. SCHRAUBEL

Problems of the research laboratory for the chemistry and biochemistry of fats in the Central Food Institute, S. IVANOV. *Chem. Umschau Fette, Oele, Wachs u. Harze* 38, 96-100(1931).—A general review of the various problems concerning the compn. of fatty acids, the formation of oils within the plant from glucose via fatty acids and glycerol, the dependence of compn. upon climatic and geographic location and the possibility of control of this compn. all with reference to the development of agriculture and the oil and fat industry in the U. S. R. R. P. ESMER.

The so-called rancidity reactions of fats. L. D. HOLDS, W. BLEYBERG, G. BRILLEY. *Allgem. Öl-Fettig.* 28, 37, 25-9(1931).—When fats contg mono- and di-glycerides are irradiated with ultra violet light in the presence of air, the free alc. groups are oxidized to aldehyde groups. Tristearin contg mono- and di-glycerides when irradiated in the presence of atm. O gave a neg. Kreis test, a pos. Ag reaction (reduction of ammoniacal AgNO_3) and a neg. peroxide test. Synthetically prep'd tristearin, free of mono- and di-glycerides, after irradiation gave neg. reactions to all 3 tests. Irradiated olive oil and also triolein gave pos. reactions to all 3 tests. Irradiated tributyrin gave a neg. Kreis test, but pos. reactions to both the Ag and peroxide tests. Before irradiation all samples gave neg. reactions to the 3 tests. Tricaprin and 2 dicaprins were synthetically prep'd, the latter from 1,2- and 1,3-dibromohydrin and were irradiated both in air and in an atm. of N . They were then subjected to modified Kreis tests and the Ag test. The Kreis test was modified by using only 0.05 g. of substance dissolved in 5 cc. benzene and in one test a 1% soln. of phloroglucinol and in the other a 1% soln. of resorcinol was used. The phloroglucinol reactions were as follows: all samples whether irradiated in air or N gave neg. results after 15 min. irradiation. The dicaprins after 4 hrs. gave a pale green and after 24 hrs. a dark green coloration. A 1:1 mixt. of tricaprin and dicaprin (prep'd. from 1,2-dibromohydrin) gave a pale green coloration after 24 hrs. The same dicaprin irradiated in N gave neg. results after 72 hrs. The reactions in the resorcinol test were essentially the same in the phloroglucinol, except that the color produced was red. The reactions in the Ag test were as follows: All samples gave neg. tests before irradiation. The dicaprins as well as the 1:1 mixt. of tricaprin and dicaprin gave pos. tests after 5 min. irradiation, while the tricaprin irradiated in air, and the dicaprin (from 1,2-dibromohydrin) irradiated in N gave neg. tests after 4 hrs. Before irradiation the tricaprin had a pleasant taste, after irradiation, an ozone like taste. The dicaprin, although they do not melt in the mouth after 8 hrs. irradiation produced a slight to distinct "scratchy" sensation. The

1 l mixt of tricaprin and dicaprin, which did melt in the mouth, also produced a "scratchy" sensation after irradiation. W. I. BOLLERS

Yellowing of the fat in Australian frozen rabbits: Its nature and cause. J. R. VICKERY *J Council Sci Ind Research* 4, No 1, 1-5(1931)—The yellowing of the fat, especially about the kidneys, is due to long-continued action of O on the unsatd fat acids, specifically linoleic glycerides. The yellow fraction of the fat was purified and resembled an oxidized acid contg 18 C atoms, the pigment existing *in situ* as an unsatd ketonic compd of undetd compn. The yellowing occurs at cold-storage temps as low as -10° , though sharp freezing greatly retards the development and the accompanying rancidity. When the fatty tissues of the rabbit were cold stored in an atm of pure N, no yellowing of the fat occurred. The yellow color appeared soon after the reintroduction of O. Various impermeable coatings were tried without success. Blood and hemoglobin are effective catalysts for the production of the yellow color; hence the complete removal of blood is advised. Some relief is obtained by packing the rabbits so that the muscular wall of the intestine covers the abdominal fat. Cold-storage temps of below -14° are recommended. C. R. FELLEAS

The detection of less than one-half per cent of oil of sesame in fats. R. LUCENRINI *Ann chim applicata* 21, 82 3(1931)—See C A 25, 2316 A. W. CONTIERI

Determination of iodine number of liquid fats with mixtures of ether and water. V. I. GSAFOVA *Zhur Prikladnoi Khimii* 3, 1207 21(1930)—The method of Margosches (C A 18, 2434) is satisfactory for liquid fats but in the presence of unsatd hydroxy acids the results are high. If not enough I_2 is added HI unites with fats. V. K.

The aniline point as a constant of fats and oils. C. G. KATRAKIS AND J. G. MEGALOKONOMOS *Praktika Akad Ithenon* 5, 207 9(1930)—To det the aniline point of fats and oils, a certain quantity of the oil or fat is filtered, neutralized with $NaOH$ (15% at 60°) and mixed with the same vol of normal benzene (Kahlbaum). Five cc of this mixt is placed in the app and 5 cc of aniline added to det the aniline point. This method is useful in detg whether or not a particular oil or fat has been adulterated by the addn of other oils or fats with the same refractive index. N. V. P.

A further study of the reactions occurring when oils and fatty acids are heated in the presence of active nickel on a support. H. I. WATERMAN AND M. J. VAN TUSSEN-BROEK *Chem Weekblad* 27, 146 7(1930) cf C A 24, 742—The hexabromide value for soy-bean oil as detd by the method of Carriere is not altered by the usual technical method of deodorizing. The hexabromide value is lowered by heating in a vacuum to 225° , the presence of infusorial earth and especially active Ni increasing the change. The acid value increases on heating in a vacuum, and the change is largest in the presence of active Ni. Elaidic acid is relatively stable toward active Ni on a support and in this respect resembles oleic acid. F. D. LERUW

Fatty-acid components of the oil from the Malayan gaval, *Tomistoma schlegelii* and some properties of the oil from *Pecten (patinopecten) yossensis* Jay. SEIICHI UENO AND NONUO KUZUI *J Soc Chem Ind Japan* 34, Suppl binding 92-3(1931)—Consts of the oil from the Malayan gaval, *Tomistoma schlegelii*, are d_{4}^{20} 0.9160, n_D^{20} 1.4623, I no 77.4, sapon no 197.6, acid no 19.4, R M no 0.54, acetyl no 22.4, unsaponifiable 0.44%. The liquid fatty acids amounted to 61.3% and consisted mostly of oleic acid with small quantities of linoleic and linolenic acids. The I no was 92.8 and the neutralization no 201. The satd acids consist of palmitic and a small quantity of stearic acid. Oil from *Pecten yossensis* Jay had the following consts d_{4}^{20} 0.9894, n_D^{20} 1.4781, I no 204, sapon no 187.7, R M no 1.2, acid no 9.5. The mixed acids + unsapon gave n_D^{20} 1.4705, I no 199.8, m p 30.9° , neutralization no 187.7, Br content 70.34%, ether insol bromides 50.46%. E. SCHERUBEL

A plant for continuous oleic acid distillation. O. KRENS *Teer u Bitumen* 29, 71-5(1931)—Crude oleic acid, the filtrate from the hydraulic fatty acid press, is heated by combustion gases to 340° in a special tubular heater. The vapors are passed through a vertical purifying column, from which any nonvaporized liquid portion flows off at the bottom to a soft pitch reservoir, while the pure distd oleic acid is condensed and gives off its heat to preheaters. P. ESCHER

Preparation and purification of linoleic acid. H. I. WATERMAN AND J. A. VAN DIJK *Verfkronek* 4, No 2, 15-9(1931)—The fatty acids of sesame oil were brominated and the tetrabromostearic acid obtained was debrominated to linoleic acid. This product n_D^{20} = 1.4681 was fractionated in a cathode vacuum. A further purification was obtained by dissolving 1 vol of linoleic acid in 2 vol of acetone (b $56.0-56.1^{\circ}$), cooling to -15° to sep out oleic acid and satd fatty acid, and then cooling to -30° to sep.

out linoleic acid, followed by refractionation in vacuum. The final product obtained had a d_{20}^{20} 0.9031, and n_D^{20} = 1.46643, n_D^{25} = 1.46944, n_D^{30} = 1.47042, n_D^{35} = 1.48228, mol refraction M_D = 59.49.

Hydrogenation of fatty oils by the so-called wet process. L. Hydrogenation in presence of nickel acetate under atmospheric pressure. SEIICHI UENO AND TAKAO YUKIMORI. *J. Soc. Chem. Ind. Japan* 34, Suppl. binding 109-11 (1931).—Catalysts were prep'd as follows: A soln. of NiSO_4 was pptd. with Na_2CO_3 and the $\text{Ni}(\text{CO})_5$ was dissolved in acetic acid and evap'd, ground and screened. (2) Ni acetate was dissolved in water and kieselguhr, equal to 4 times the wt. of Ni in the acetate, was added, this was then dried and screened. (3) Kahlbaum Ni acetate was crushed and screened in the same manner. (4) Kahlbaum Ni acetate and kieselguhr were mixed as above. (5) Kahlbaum Ni acetate and pumice were mixed as in (2). (6) $\text{Ni}(\text{CO})_5$ was mixed with kieselguhr equal to 4 times the wt. of Ni in the carbonate. The oil used had an I no. 188.3 and acid value of 0.5. Catalysts from 0.2 to 0.6% were added to 40 g. of oil and the temp. was raised and held 5 min. at 230° and then lowered to 180° and hydrogenation carried out for 1 hr. Catalyst No. 6 was first reduced in H_2 . It was very difficult to control the expts. by this method and app. The rate of decrease of I nos. was not proportional to the amt. of catalyst used, and the catalytic power was far inferior to that of catalyst no. 6. E. SCHERUBEL.

Absorption of oxygen by unsaturated oils. I. Influence of solvent on the rate of absorption. SEIICHI UENO, ZENSABU OKAMURA AND TEIJI SAIDA. *J. Soc. Chem. Ind. Japan* 34, Suppl. binding 106-8 (1931).—The accelerating effects of various solvents on the oxidation of unsat'd oils was studied. A shaking app. was used. Co linoleate dimer was added to linseed and sardine oil. The rate of oxidation was est'd from the vol. of absorbed O_2 . A comparison of the effects of various solvents on the rate of oxidation shows in decreasing order for linseed oil as follows: glacial AcOH , decalin, Et_2O , turpentine, CHCl_3 , acetone. With sardine oil: glacial AcOH , decalin, turpentine, acetone, CHCl_3 , petr. ether. F. SCHERUBEL.

Hydrogenation of oils at high temperatures and under high pressures. L. Hydrogenation with nickel under constant high pressures. SEIICHI UENO, TAKAO YUKIMORI, HIROKO TSUCHIKAWA AND SEICHO UENO. *J. Soc. Chem. Ind. Japan* 34, Suppl. binding 111-5 (1931).—Three hundred g. of sardine oil of I no. 171.0, sapon no. 190.7 and acid no. 1.05 was used with 3 g. catalyst consisting of 22.5% Ni and 77.5% kieselguhr in a steel tube of 2 l. capacity. The expts. were carried out with 50, 40, 30, 20 and 10 atms for 1 hr. at temps. of 150° , 180° , 200° and 230° . The tube was shaken by a Slits's shaking mechanism. The following data were made upon the hardened product in p. I no., sapon no., thiocyanogen no. The mixed and solid fatty acids were analyzed for m. p., I no. and neutralization value. The reaction was most rapid at temps. of 180° and 200° . Results are tabulated. F. SCHERUBEL.

Solidifying point of vegetable oils. J. MÓRUSZ. *Kisérlet Közlemények* 31, 321-31 (1931).—Good results were obtained by a modified Mohr method with 6-7 g. of oil instead of 30-35 g. Solidifying points do not increase proportionally with addn. of solid triglycerides. Solidifying point curves bear a relation to the consistency of solid fats. The solidified fat is cryst. if the temp. increases upon solidification. Solidification curves of some glycerides change upon adding other fats to them, while some retain their original characteristics. S. S. DE FINLEY.

The cottonseed-oil industry. E. L. CARPENTER AND LEO HOLDREDS. *Mech. Eng.* 53, 353-9 (1931).—Its history, economics, processes and problems. E. H.

Preliminary study on peanut varieties at the Lamsa Experiment Station, Lamsa, Batavia. PEDRO L. PAULINO. *Philippine J. Agr.* 1, 273-89 (1930).—The oil contents of 13 varieties of peanuts are given. Varieties proving rich in oil, as per sample submitted, were Japanese, 58.90%; Tennessee Red, 56.40%; North Carolina Runner, 54.37% and the San Jose No. 1, 54.13%. Other varieties had 44.99-54.10% oil content. By analysis on the dry basis Tennessee Red contained 60.07% of oil, Japanese, 58.80%, San Jose No. 1, 58.08%, and San Jose No. 2, 53.01%; while other varieties contained 49.08-56.94%. JERRY O. HARDESTY.

Greek tobacco-seed oil. J. D. KANELIS AND N. S. KARMIS. *Praktika (Abd. Athina)* 4, 475-81 (1929).—The (drying) oil (36.85%) had: I value 117.8-137.9, dn 0.9253-0.9440, n_D^{20} 1.4735-1.4828, acid value 2.25-16.93, sapon value 186.6-201.6, Hehner value 95.21-96.34, Reichert Meissl value 0.32-2.03, Polenske value 0.15-0.30. B. C. A.

Factors of quality in copra. F. C. COOPER. *Malayan Agr. J.* 19, 128-36 (1931).—About 6% represents the water content of the av. stable condition for kiln-dried

copra stored in bulk. Higher moisture produces deterioration with development of heat, acidity, rancidity and internal darkening and formation of white, brown, green and black molds. As deterioration progresses the oil % increases because of the disappearance of tissue in the form of gas and water vapor. It has been found that bad Malaya copra with a high degree of rancidity, if stored under free ventilation, will diminish in acidity, while the copra, as a whole, will become dry and clean and the mold will fall away as dust. Drying arrests acid formation from the neutral oil, while the existing fatty acids are converted into water vapor and gas and thus reduce the free acid. Rubbery copra does not press below 7% oil content; it results from young nuts. Copra can be compressed to advantage for shipment at a pressure of 15 cwt. per sq. in. without loss of oil. E. SCHREIBER.

Examination of textile oils. I. RALPH B. TRUESLER. *Oil and Fat Ind.* 8, 51-7, 77 (1931).—The app. that was designed to measure the susceptibility of an oil to oxidation consists of 3 parts: a reaction bulb, a gas-measuring device and a shaker. The bulb is made of pyrex glass and is 134 mm. long and 40 mm. in diam. and holds 130 cc., which accommodates 25 g. oil and 100 ml. of O. An inlet for O is provided by a capillary tube with a stopcock. The oil sample is admitted through a sep. capillary tube near the bottom of the bulb. The gas measuring device consists of a leveling flask of 300 ml. capacity contg. sufficient Hg to fill the tubing and to maintain a column in the stem of the manometer and to fill the gas buret. The flask is adjustable by means of a ring clamp. The balancing manometer serves as a means for attaining a standard pressure in the buret when a vol. of O is to be detd. It is made from a 100 ml. pipet. The shaking device provides the means for oxidizing oils under conditions that can be controlled and duplicated. It consists of a rocker arm 14 mm. long from the axis of the shaft to the center of the bulb. The upper part of the rocker arm from the axis of the shaft to the pivot of the rocker beam is 7.5 cm. The rocker beam is given a transverse motion by a rotating wheel about which it travels in a circle of 36 mm. diam. at its point of attachment. II. *Ibid.* 103-6, 109.—A measure of the oxidizability of oils to predetermine their desirability for textile purposes can be ascertained by treating 25 g. of oil with a definite amt. of O in a reaction bulb at a definite temp. in an oil bath and shaking at the rate of 140 shakes per min. for 1 hr. in a shaker. It was found that fatty oils appear to pass through 3 stages of oxidation. At ordinary temps. peroxide combinations are produced which are assocd. with the subsequent occurrence of rancidity. Increased temp. during oxidation causes stable oxygenated compds. to be formed. At higher temps. oxidation resembles combustion in that CO₂ is evolved. This reaction is most common with unsatd. acids. The accuracy of the method is approx. 0.2%. Dry oils should be used. The new method permits numerous applications to problems dealing with the behavior of oils toward oxidation. E. g., the presence of unsatd. free acids in a glyceride was found to render an oil more sensitive to oxidation. The accuracy of this test permitted a study of textile oils regarding their stability toward O and also gave a means for evaluating antioxidants and for measuring their effect upon oxidizable oils. E. SCHREIBER.

Toilet-soap manufacture. W. H. SIMMONS. *Mfg. Chemist* 2, 95-9, 107 (1931).—A review. E. H.

Unsaponifiable matter and so-called iso-oleic acids contained in toilet soaps on the market. SEI-ICHI UENO, GENTARO INAGAKI AND HEIJIRO TSUCHIKAWA. *J. Soc. Chem. Ind. Japan* 34, Suppl. binding 90-2 (1931).—The unsaponifiable matter of 12 samples of soap was detd. by Spitz and Hömig's method. The insol. mixed acids were subjected to the Pb salt alc. method for sepn. into solid and liquid acids. The iso-oleic acids were calcd. by A. S. Richardson's formula and varied from 2.2 to 5.5%. It has been shown by the authors that 36-43% of the total acids are converted to iso-oleic acids during hydrogenation and by calcn. about 5 to 10% of hardened oil was used in the raw material for these soaps, some of which were toilet soaps. E. S.

Influence of the component glycerides of the soap base in the production of a white (milled) toilet soap when dried in a Procter dryer. G. F. CALVEY. *Chem. Eng. Min. Rev.* 22, 472-3 (1930).—The transparent appearance, characteristic of milled soaps made from soap that has been prepd. by sudden chilling in a cooler and dryer, can be avoided by careful control of the proportions of solid and liquid acids in the soap stock. B. C. A.

Soap solutions. IX. Surface tension of aqueous solutions of binary mixtures of soaps having extremely different molecular weights. J. MIYUKO. *J. Soc. Chem. Ind. (Japan)* 34, Suppl. binding 116-6 (1931). cf. *C. A.* 25, 427.—Na salts of fatty acids whose mol. wts. are very small or very large are inefficient as detergents. Surface tension at 40° was measured by Traube's stalagmometer for aq. solns. of binary mixts.

of the lowest Na soaps of said acids C_8 to C_{18} and C_{11} and also Na erucate, C_{22} . In these systems this soap is the primary adsorbate and dominant in quantity at the surface, even in the presence of more capillary active laurate, C_{12} . Soaps of C_{12} to C_{22} are not able to improve the surface activities of their solns by the addn. of soaps below C_{12} . Sometimes soaps below C_8 have a coagulating tendency upon the C_{12} - C_{22} soaps because they are common electrolytes in dil. solns. J. SCHREIBER.

Migration of sodium chloride in soap cakes. J. MIKUMA. *J. Soc. Chem. Ind. (Japan)* 34, Suppl. binding 116(1931).—The fact of salt migration was verified for all aged cakes of boiled and settled and even milled soaps. The direction of migration is detd. by the difference in vapor tension between the outside and inside of the cake. M. believes that NaCl and other sol. impurities diffuse through the interfibrillar water of the cake mass. Sepd. curd soaps vary their structure during cooling, maturing and storage, becoming more and more fibrous, liberating a part of the hydration H_2O as free (syneresis). Tenacity of the soap base seems to be closely related to the nature and content of soaps of higher liquid acids (oleate, ricinoleate, linoleate) which form the basal structure by cementation of total mass and also give finer fibers to soaps of higher solid acids which otherwise would be very hard and brittle. F. S.

Rayon oils (MULLIN, CALDWELL) 23. The Coral pea or *Adenarthera parsonsia* (PIERARIS *et al.*) 11D. Oil filters (Fr. pat. 697,949) 1. Oxidizing paraffins, etc. (Ger. pat. 522,361) 10.

Fats and waxes. JOHN G. STEPHENS. Fr. 697,572, June 17, 1930. Fats and waxes are atomized and coated with a harmless powder such as flour to prevent agglomeration and keep them in an easily used form, e. g., as ingredients for cakes, confectionery, etc.

Saponifying fats and oils. CLEMENS BERGELL. U. S. 1,799,493, April 7. In soap manuf. comprising a first and main saponification of oils and fats succeeded by a second and final saponification, the first sapon. is effected by boiling with an alkali lye, a small quantity of an electrolyte and water (the mixt. being boiled contg. about 55% of combined and uncombined fatty acids and 10-1.5% of the electrolyte) so that the product after boiling can be rendered viscous by addn. of a small proportion of water, and the second sapon. is effected by boiling after adding sufficient water that the mixt. contains about 80% of combined and uncombined fatty acids.

Soft fat or oil from solid fats. YOSITORA IWAMOTO. Japan. 90,175, Feb. 6, 1931. Soft fat or liquid oil is manufd. from solid fats by the action of silent discharge against the fats in *vacuo* or in N_2 or H_2 .

Rotary drum apparatus suitable for extracting animal or vegetable oils or fats with solvents or for other extraction operations. WALTER E. SANGER and OSCAR H. WEXSTER (Sanger's executor to Wuxster). U. S. 1,799,481, April 7. Structural features.

Polymerizing oils and fats. I. G. FARBENIND A-G. Fr. 697,785, June 23, 1930. Products particularly for use in the textile, linoleum and rubber industries or as addns. to varnishes, etc., are made by polymerizing vegetable oils or fats or their derivs. in the form of emulsions, small quantities of O or S or compds. liberating them being used as accelerators.

Apparatus for extracting oily substances from grain, etc. PHILIPP L. FAUTH. G. M. B. H. Fr. 697,206, June 12, 1930.

Apparatus for removing oil from soy-beans and for semi-baking the beans. TEIKICHI SATOH. U. S. 1,799,256, April 7. Various details are described of a rotatable drum app. which may be supplied with solvent.

Extracting oily substances from citrus fruit peel. WILBUR A. PIPKIN. U. S. 1,798,555, March 31. App. and mech. details of operation are described.

Wax compositions. I. G. FARBENIND A-G. Fr. 698,643, July 8, 1930. Wax compns. contain wax and org. derivs. of NH_2 contg. at least one alkoxy or cycloalkoxy radical or both, or salts or derivs. of these compds., e. g., mono-, di- or tri-ethanolamine. Cl. C. A. 24, 1535.

Cleansing agent. MAURICE SADO. Fr. 697,751, June 21, 1930. A cleansing agent is composed of black soap, powd. $CaCO_3$, boric acid, Na_2CO_3 , and water.

"Washing compound." JOSEPH A. SEIBERT. U. S. 1,799,930, April 7. Coloring matter, perfume and alc. are successively added to a mixt. of Na_3PO_4 , powd. soap and borax, with thorough mixing.

Soap. SACCHARIN FABRIK A-G. NORM. FAHLBERG LIST & Co. Ger. 522,367, Apr. 22, 1928. In the manuf. of soap contg. salts of aromatic chlorosulfonamides, decomposition of the salts in the presence of the soap is avoided by preliminarily treating the

snap or the fat from which it is prepared, with hypochlorites or Cl in an amt. exceeding that necessary for bleaching.

Soaps, etc. *Soc. Parf. & Cos.* *Fr. P. 8716*, (Oct. 11, 1929) — Soaps of all kinds, creams, etc., are improved by incorporating therewith veg. materials from the secretions of glands of the digestive organs of an animal, particularly the intestines.

Soap powder. *Adolf Welter*. *Ger. 516,913*, Nov. 1, 1928. See Int. 372,740 (C. A. 25, 490).

28—SUGAR, STARCH AND GUMS

J. K. DALP

The sugar plant of the future. *FORWARD GUYWITT*. *Listy Cukrovar*, 49, 265-6 (1931) — A discussion of allied industries which may become subordinated to sugar plants as artificial leather, silk, gums, adhesives, etc. *FRANK MARPISH*

A review of Czechoslovakian patents pertaining to sugar technology. *Jiří Vondrák*. *Listy Cukrovar* 49, 299 2 (1931) — All patents are listed and described briefly. *FRANK MARPISH*

A contemporary comparison of the sugar-cane and sugar-beet industries. (1) *FALLANA*. *Chem. Listy* 25, 111-4 (1931). *FRANK MARPISH*

The revaluation of raw sugar types used by the Prague exchange for goods and valuable papers. *K. ŠAVORÁK*. *Listy Cukrovar* 49, 267 71 (1931) — On account of the effect of sunlight etc. upon raw sugars the standards have to be rechecked every 2 yrs. The color of the specimens is detd. relative to one of the standards of a medium color. After a search for mineral standards, powd. glass resembling sugar in color and in grain size was chosen as a standard. Complete matching was not possible, for sucrose is cryst. optically anisotropic and colored only on the surfaces; glass has lustrous cleavage, is optically isotropic and is colored throughout its thickness. All standard types of raw sugar showed an increase of 0.04 g. in wt. during the 2 yr. period as a result of an increased humidity. In 31 of the standards, the change in absorption did not exceed 1.5%, 6 standards showed an increase greater than 3%. 2 were altered to visual inspection and discarded. Since all standards are kept in the dark at 20° S. thinks that changes in humidity lead to a spontaneous refining of the raw sugar. The small changes permit the same standards to be used for another (2nd) year. Frequent (monthly) retesting is of no value. *FRANK MARPISH*

Adsorption from sucrose solutions. IV. Sucrose according to the stratum method. *J. VASÁTKO AND J. VÁCHA*. *Listy Cukrovar* 49, 215 9 (1931), cf. C. A. 24, 6053 — Sucrose solns. lower the decolorizing effect of activated charcoal. Sucrose solns. (d. 1.0904, ρ_n 7.3) at 20° were passed through 10 g. dry charcoal spread out in layers 7-8 mm. thick. The vol. of sucrose passing through the charcoal layers increases linearly with time. The time required for the flow of 50 cc. of the above sucrose soln. varied with different charcoals: standard norite 27 sec., supranorite (5+) 41 sec., polycarbon 49 sec., supranorite (3+) 53 sec., carboraffin 56 sec., supracarboraffin 62 sec., and radite 93 sec. The same values were obtained with water, in dil. solns., sucrose has little effect upon the rate of flow through charcoals. The dry charcoal adsorbs more water than sucrose during early filtration. Plotting the adsorbed sucrose against vol. of soln. filtered showed a typical isotherm with an asymptotic approach at a vol. of 50 cc. of filtered soln. Conc. sucrose solns. deviated greatly from water curves, during the start of filtration water is adsorbed more rapidly than sucrose so that even a neg. absorption of sucrose may occur. The max. sucrose adsorbed by this layer method is the same as the max. sucrose adsorbed by charcoal in suspension, providing the same ratio of soln. to charcoal exists. Leaching expts. show a complete reversibility of the adsorption of sucrose upon charcoal. *FRANK MARPISH*

ρ_n and acidity determinations of cane juices expressed by a fourteen-roller milling plant. *FRANCISCO S. GOMEZ*. *Philippine Agr.* 19, 609 34 (1931) — Detns. of total acidity by titration, ρ_n by quinhydrone electrode and Brix solids were made of the juices coming from each mill of a sugar central. The total acidity of the first mill juice is greater than that of crusher juice. This is not due to bacterial action but to expression of acids from the rind of the cane. Large amts. of sucrose are lost when juice is allowed to stand more than 24 hrs. When the mills are thoroughly washed once every shift, bagasse is periodically removed and the juices are handled promptly, the increase in acidity due to bacteria is negligible. The use of chemicals in cleaning the mills is not advisable. *A. L. MEHRING*

The results of eugenic experiments with sugar beets. F. CHMELAK AND JAR. SIMON. *Listy Cukrovar* 49, 263-70(1931)—Four domestic varieties of sugar beets, 4 German and 1 Polish were grown on 12 plots. The order of total yield of sugar was Dobrovická, Zapotilova, Selecta Praha, Wobanka, Schobler, Schreiber SS, Kleinwanzleben N Mette and Buszczynski P. Dobrovická, Zapotilova and Selecta Praha were characterized by a high yield and good sugar content and excelled all foreign varieties in total sugar yield. Wobanka had the highest sugar content but only an average yield. FRANK MARSH

The increased sugar-beet yield by early planting and late harvesting. KOCH. *Zuckerzudenbau* 12, 73(1930), *Listy Cukrovar* 48, No 12, Rozhledy 41—Late harvesting is recommended. FRANK MARSH

The time for beet thinning. KAMLAH. *Zuckerzudenbau* 12, 78(1930), *Listy Cukrovar* 49, No 4, Rozhledy 1. FRANK MARSH

The prediction of the quality of a seasonal run of beets on the basis of beet analysis. JIKL VONDRÁK. *Listy Cukrovar* 49, 333-7(1931)—The quality of the alkali of beet liquors was predicted correctly in 10 out of 12 cases from the amide N content of the beet. V considers the amide N an index of what the alkali of the juices will be from the same beets unless the alkali is not decreased by the inversion of sucrose. The frequency of rainfall is the most reliable criterion for making the above prediction. F M

Parasites and diseases of beets for 1930 (Czechoslovakia). F. RAMBOUSEK. *Listy Cukrovar* 49, 357-57(1931)—A detailed annual report is given for all forms of parasites including the extent, no., type of destruction and effective means of eradication. Beet pathology is discussed in the same manner. FRANK MARSH

A negative correlation observed between the nitrate nitrogen in the juice of beet leaves and the weight of the leaves. DONALD E. FAFAR. *J Agr Research* 42, 53-6(1931)—In order to reduce the possibility of serious sampling error it is necessary to choose leaves that are approx. the same size. W H ROSS

The yield of sugar from dried slices. J HAMOUS. *Listy Cukrovar* 49, 352-3(1931)—A computation is given showing the cost of manufacture of sugar from dried beets. FRANK MARSH

The increase of polarization of sugar-beet mash by grinding, mixing and drying. O KORECKÝ. *Listy Cukrovar* 49, 275-83(1931), *et C A* 24, 312—Samples ranging from 100 to 1700 g were ground, mixed and weighed. The following facts were observed: (1) Beets lose weight during weighing. (2) Beet mash loses weight during mixing and weighing. (3) The loss in wt. is due to an evaporation of water, it is responsible for higher polarization values for mash than for beets. (4) The loss in wt. and increase in the polarization vary directly with the time of grinding, time and rate of mixing, time of analysis, temp. and humidity of the atm. and surface area of the exposed mash. (5) The smallest loss in wt. occurs during the weighing, the largest during grinding. (6) The increase in polarization always exceeds 0.1%, and may exceed 0.2%. F M

Experiments on decolorizing greens from refined loaf sugar with norite. I F ZELDMAN AND A I SICHKAR'NEKO. *Nauk Zapiski Trukrom Prom* 10, 403-17(1930)—Greens of 65° Brix mixed with 0.75% of standard norite were heated for 15 min. at 60°, 75° and 90°. The increase of the temp. from 60° to 75° gave a decrease in color of about 7% and with a further increase of the temp. to 90°, the color decreased about 5% more. To prevent an increase of reducing sugars at the high temp. all further expts. were made at 75°. Decrease of the d from 70° to 50° Brix gave about 20% more decolorizing effect. An increase of the time of contact of norite with greens from 15 to 45 min. gave a very slight increase in decolorization. An increase of standard norite from 0.6 to 1.0% gave a noticeable decrease in color. A further increase of norite to 1.25% did not show any further benefit. Expts. at the refinery gave better results than in the lab., probably because of the thickness of the norite cakes in the filter presses. The quality of the sugar from greens treated with norite was lower than that of refined sugar of middle quality. Norite was regenerated by the following scheme: Discharged from filter presses it was mixed with water to about 20° Bx and boiled with HCl for 1 hr., filter pressed and washed to a slight acid reaction. It was then heated in retorts. Norite used for lower grades of refined sugar was first boiled with NaOH for 2 hrs., then treated as above. Regenerated norite gave very poor results in decolorizing greens. The decolorizing power of norite depends on the initial color of the liquors. Hence decolorizing of refinery greens with bone black is more advisable as the decolorizing results are better and manufacturing costs are lower. V E BAIKOW

The Teatin method of clearing liquors. K. SMOLEŃSKI AND M. WERKENTHIN-GWNA. *Gazeta Cukrownicza* 68, 145-8(1930), *Listy Cukrovar* 49, No 26, Rozhledy 24—The process of using 1.01% CaO in the Hougaerde factory decreased the time of

filtration and extra. of sludge with a high CaO recovery. The elements of the process throughout the process are given.

The course of the first saturation. V. Saturation with oxalic acid in solution of Ene and sucrose. J. DENEK AND D. IVANKOV. *Listy Cukrovar.* 49, 271-2 (1931), d. C. A. 25, 1700-1.—Studies in the course of sats. of a CaO mixture with alk. (14) showed that the origin and quantity of sediment are not proportional to the quantity of CO_2 , and that with a delayed or sudden formation of sediment, CaO and sucrose are removed from soln. Further sats. with CO_2 allow both CaO and sucrose to reenter the soln. from the sediment. The release of CaO exceeds the rate of sats. with CO_2 and leads to a temporary increase in alk. Solns. contg. 15% sucrose and 1% CaO were treated with increasing quantities of $(\text{COOH})_2$. The course of the first sats. with $(\text{COOH})_2$ gave the same picture of changes in alk. as CO_2 , the withdrawal of CaO from soln. by the sediment occurs only at full sats. or at elevated temp. The addn. of CaC_2O_4 to CaO-sucrose-solns. causes an immediate withdrawal of CaO from soln. with the sediment. The alk. of the solns. after equil. is reached corresponds to a satd. CaO soln. CaO from a freshly or undersatd. soln. cannot be adsorbed by sediment. Sucrose is only adsorbed with adsorbed CaO. A max. of 2.5 equivs. of CaO is ptd. by CaC_2O_4 .

FRANK MARSH

Glycerophosphoric acid during filtration and saturation. BLANOSLAY DROBKY. *Listy Cukrovar.* 49, 371-4 (1931)—Sucrose solns. (15%) contg. 0.2 or 2.0% CaO were treated with Ca glycerophosphate (0.3 g. per 100 cc.) and digested 12-48 hrs. at 20° and 100°. The amt. of Ca glycerophosphate hydrolyzed ranged from 2.6 to 4.8%; and independent of temp. and time. During the sats. of a 15% sucrose soln. contg. 2% CaO at 80°, the hydrolyzed Ca glycerophosphate did not exceed 6.3% and 63.5% of the phosphate was carried down with the sediment in an insol. form. The ratio of inorg. phosphate to org. phosphate was 1:13. In com. prepn., the reverse ratio holds because of the presence of inorg. phosphate in the root of the beet. During the 2nd sats., only traces (0.005%) of the Ca glycerophosphate hydrolyzed and 6.2% was ptd. with the sediment in an insol. form. The total Ca glycerophosphate recovered in analysis was 96.4%. Solns. contg. 15% sucrose, and 1.2401 g. Ca glycerophosphate per 100 cc. were made slightly acid with HCl and treated with 3.5 g. CaH_2O_4 . The ratio of org. P to inorg. P was 1:5. After heating the solns. were made alk. with milk of lime, cooled, made up to 1 l. vol. and filtered. The Ca phosphate ppt. carried with it 65.1% of the added glycerophosphate. The addn. of 1.4 g. CaH_2O_4 to the preceding clear filtrate and a repetition of the treatment carried down 82% more of the org. P. The Ca glycerophosphate is hydrolyzed very little to glycerol + Ca phosphate during filtration, and sats.; it is deposited without decompt. in the sediment. This ptn. is due to the adsorption of the org. P upon the ready formed Ca phosphate ppts. Further studies with glycerophosphates from leucithin are promised.

FRANK MARSH

The porosity of sediments. J. DENEK AND J. DOSTÁL. *Listy Cukrovar.* 49, 375-82 (1931)—A series of runs with Knoop, Sweetland and Kelly presses contg. sats. sediments is described in detail. Increasing addns. of CaO up to sats. decrease the porosity; in a semi-com. plant when the CaO was increased from 1.2 to 4% the porosity decreased from 62.4 to 43.8. This is a reversal of the experience in com. operation. D. ascribes the discrepancy to an increased sp. gr. of the sediment with a decreased vol. of porosity.

FRANK MARSH

Antiseptic medium "Lystonol." J. ZALESEK. *Gazeta Cukrovarnicka* 66, 653-5 (1930). *Listy Cukrovar.* 49, No. 4, Rozhledy 4.—During the foaming of liquors in the diffusion process, 20-40 g. Lystonol is added per ton of beets to prevent infection. Analysis shows H_2O 72.43, HCHO 0.79, ZnCl_2 8.03, NaCl 17.07, CaSO_4 0.46, FeCl_3 0.27, and NH_4Cl 0.56%. The active disinfectants are ZnCl_2 and HCHO .

F. M.

Volumeter of modified construction and its manipulation. V. YANOVSEK. *Nové Zprávy Cukrovarnické Prom.* 10, 423-9 (1930).—The vol. of a lump of sugar is detd. by the difference in the vols. of kerosene in a volumeter (C. A. 23, 297; 24, 257). Instead of kerosene Hg can be used and has the advantage that the same sample of sugar can be used for crushing and soly. tests. The methods for the detn. of the constant of the app. and limits of the error are described.

V. E. BAIKOW

Isolation of choline from beet molasses. V. STANEK. *Z. Zuckerind. Czechoslov.* 55, 103-5 (1931)—See C. A. 25, 1406.

G. G.

Slow-boiling massecuites. S. A. WICKBY. *Intern. Sugar J.* 32, 594-5 (1930).—See C. A. 25, 1114.

E. J. C.

Seedling low-grade massecuites. G. F. MURRAY. *Intern. Sugar J.* 32, 595-6 (1930).—See C. A. 25, 1114.

E. H.

Determination of total solids in massecuites, drainage liquors and molasses.

SAILLARD AND DARTOIS. *Suppl. etc. hebdomadaire central fabr. sucre* 42, No 2168 (Oct., 1930), *Chimie & Industrie* 25, 440(1931).—As the usual method requires 5-6 hrs., Morizot has proposed a quicker method consisting in absorbing 10 cc of a 20% soln of masecuite in absorbent cotton contained in a glass tube 6-cm long and 3 cm in diam and passing heated air through the tube, one end of which is open and the other connected to a suction pump. The tube is placed in a double wall oven, heated by oil to 170°. At this temp and with a vacuum of 43 cm, drying is complete in 2.5 hrs.; it can be shortened by increasing the temp and the vacuum. A. PATTINAU COULTRE.

The effect of N fertilizers on the yield and quality of sugar beets (MIKHAILOVSKII) 15. Heart rot of sugar beets (GARMANN) 15. Cellulose amyloid (JUSTIN MUELLER) 23. Soil and cane composition in relation to Lahaina failure at Waipio Substation (HANSSON) 15. The measurement of cuprous oxide with permanganate solution (BAUTINS) 7. Recent investigations on the constitution of pectins [sugar beet] (BAIDEL) 10. Adsorption of sugars by animal charcoal and vegetable decolorizing carbons (TAKETOMI) 2. The effect of mineral fertilizers and manure upon the quality of sugar beets (SIRACHENKO) 15. The constancy of vegetable composition according to Lieng and the sugar beet bred by selections (SAILLARD) 11D. Sugar cane α -cellulose (DE LA ROZA) 23. Fertilizer from sugar factory residues (U S pat 1,709,170) 15. Drier for beets (RUSS pat 19 535) 1.

Sugar purification. MANUFACTURE DE PRODUITS CHIMIQUES DE JOU-EN-JOAS (ANCIENS ÉTABLISSEMENTS LOUIS DESCAMPS) Fr 1,077,824, June 23, 1930. A colloidal suspension of a basic Zn hyposulfite, aldehyde or not, or like hyposulfites is prep'd by adding to the hyposulfite the desired amt. of an earth or alk. earth base corresponding or not to the metal of the hyposulfite. The products are used in sugar or pectin manu. as flocculating agents for colloidal impurities.

Purifying sugar solutions. HOLZHYDROLISE A-G. Ger 522,360, Aug 6, 1929. In the electrosomotic purification of sugar solns obtained by hydrolyzing cellulose, starch, etc., with acids, a three-cell app. is used in which the anode and cathode compartments contain a sugar soln of approx. the same concn. as the soln to be purified, the latter being contained in the middle compartment.

Diffusion apparatus for extracting sugar, etc. AUGUSTE FUGÈVE VASSEUX. Fr 698 191, June 28, 1930.

Apparatus and method for disintegrating and extracting juice from sugar-cane stalks. WM H MORGAN, SR (to Morgan Hurrycane Co.) U S reissue 18,022, March 31. Reissue of original pat No 1 646,762 (C. A. 22, 329). The stalks are reduced to a loose fibrous mass by longitudinal splitting.

Depurating a mixture of beets and water. ARNALDO GUADAGNINI. U S 1,798,792, March 31. A mixt. of beets and water for the production of a saccharine liquid is heated in a sealed vessel to effect cooking under a pressure of about a half atm., and vapor in the upper part of the vessel is superheated by a steam coil for quickly blowing out the vaporized impurities from the vessel, the mixt. is subjected to a further cooking operation under pressure, and the vapor in the upper part of the vessel is again superheated for blowing off the impurities which are vaporized. App. is described.

Apparatus for making an improved caramel. FUGÈVE BASSIÈRES. Fr 697,998, Mar 25, 1930.

Apparatus for liming molasses, etc. CARL STEFFEN, JR. Ger 522,359, Dec 19, 1925.

Washing apparatus for separating starch from starch-containing materials. OTTO RIEMANN. U S 1,798 811, March 31. Structural features.

Adhesive gum. MARIO VISTAKINI. Fr 698,074, June 26, 1930. An adhesive gum in powder form is obtained by adding powd. gum tragacanth to a mixt. of salicylic acid and boric acid.

29—LEATHER AND GLUE

ALLEN ROGERS

Various vegetable tanning materials from Eritrea. G A BRAVO. *Ann. chim. applicata* 21, 141-6(1931).—Several plants have been found in Italian Eritrea which contain high percentages of tanning materials. Among these *Oxyris abyssinica* (Fritrea sumach) and *Acacia nilotica* both have considerable tannins, 18.4 and 35% resp., the first is widely diffused, but is readily affected by parasites while the latter is compara-

tively rare. However, *Coccolopia tinctoria*, having tannin contents up to 51%, has been cultivated, and promises to become a source of tannins, and also to become a good substitute for sumach. A. W. COCHRAN

Analysis of miscellaneous tannery materials. Committee report. V. J. MERRILL. *J. Am. Leather Chem. Assoc.* 26, 244-57 (1931).—Provisional procedures (based on familiar methods) are submitted for the routine analysis of sodium sulfide, lime, 1 room salt, soda ash, NaHCO_3 , NaOH , H_2SO_4 , and NaHSO_4 . H. B. MERRILL

Determination of moisture in tanning extracts. Committee report. 1930-31. GENE W. STANBERRY, et al. *J. Am. Leather Chem. Assoc.* 26, 244-6 (1931), cf. *C. A.* 23, 3368.—Samples of quebracho and chestnut exts. were analyzed for H_2O by the direct and by the indirect methods. Quebracho showed no difference outside the permissible range of exptl. error. Chestnut gave lower results by the direct method. Lowest total solids resulted from 6-16 hrs. drying in the direct, and 16 hrs. in the indirect method. Variations in humidity are without effect when the prescribed dryer is used. The electric oven at 99-103° gave results for H_2O that are very slightly higher than those obtained with the "evaporator dryer." H. B. MERRILL

Quantitative method for determining tannin colorimetrically and the value of the method of P. Menaul. A. B. SHAKHILDIAN. *Zhur. Prikladnoi Khim.* 3, 1117-24 (1931).—The unknown soln. should contain 0.5-2.0 g. tannin per l. Add to 100 cc. of it 0.315 g. chem. pure tannin dried at 100° and to 10 cc. of the resulting soln. add 10 cc. of a 2% $(\text{NH}_4)_2\text{MoO}_7$ soln. and dil. to 100 cc. To 10 cc. of the unknown soln. add also 10 cc. of a 2% $(\text{NH}_4)_2\text{MoO}_7$ soln. (neutral salt can be used) and dil. to 100 cc. After 15 min. standing compare both solns. in a Duboscq. colorimeter and calc. the results, assuming that 0.315 g. chem. pure tannin of Kahlbaum contains 0.015 g. impurities. This method is only $1/10$ as sensitive as the method of Menaul (*C. A.* 18, 1005) but is simple, rapid and the color produced is more stable. V. KALICHINSKY

Some notes on microscopic technique with special reference to microtannology. FRED O'FLAHERTY. *J. Am. Leather Chem. Assoc.* 26, 257-64 (1931). H. B. M.

Tanning and dyeing of furs. A. DOMINGUE. *Boll. ufficiale staz. sper. ind. pell. mat. conciansi* 8, 168-71 (1930).—Directions are given. G. SCHWACH

Application of colloidal mulls to the leather industry. ALAN A. CLAFIN. *Leather Manuf.* 42, 59-60 (1931).—It is suggested that better emulsions with less emulsifying agent can be made for fat liquoring by the use of the colloid mill. H. B. MERRILL

Skin color tests. Committee report, 1931. C. A. BLAIR, et al. *J. Am. Leather Chem. Assoc.* 26, 223-34 (1931), cf. *C. A.* 25, 2325.—Pickled calf skin and sheep skivers were depickled with borax, adjusted to pH 4.8 with acetate buffers, and preserved with borophenol soln. Tanning tests were made with cutch, quebracho and chestnut exts. under a variety of conditions. Collaborators' comments indicate that the method is generally satisfactory, divergent opinions are expressed regarding working details. H. B. MERRILL

Collagen and gelatin. G. C. HERING AND H. R. KRYT. *Chem. Weekblad* 28, 142-50 (1931).—A review. F. DE LEEUW

Extracts obtained in the manufacture of etheral oils from conifers (BOBROY) 22. Röntgen diagram of collagen (HERZOG, JANCKE) 2.

Synthetic tanning agent. J. R. GREGY A-G. Ger. 515,664, Aug. 17, 1926. A tanning agent free from mineral acid is prepd. by treating neutral salts of the tanning material from aromatic hydroxysulfonic acids and CH_3O (the soly. of which in water depends on the presence of the HSO_3 group) with salts of H_2SiF_6 . Thus, the tanning agent from cresolsulfonic acids and CH_3O is neutralized with Na_2CO_3 and dried. Na_2SiF_6 is then added. A further example is given.

Tanning green hides. JOSEPH M. BROWN. U. S. 1,800,131, April 7. The de-haired hides are treated with an aq. soln. of gum catechu ext., further subjected to the action of the soln. with addn. of fluid ext. of wild haw and alum, then washed and dried.

Drying apparatus for hides and furs. A. A. POPIYADUKIN. Russ. 19,534, Feb. 28, 1931. Constructional details.

Leather. PAUL RAKSANYI. Austrian 121,993, Nov. 15, 1930. Non-slipping chrome leather capable of being nailed is manufd. by steeping the prepd. hides successively in 3 baths contg. (1) water glass 25 and linseed oil 2%, (2) CaCl_2 10 and KMnO_4 1%, and (3) chrome alum 20 and FeSO_4 2%. The proportions stated may be varied, and CaCl_2 in the second bath may be replaced by equiv. salts.

Greasing leather. HANSEATISCHE MÜHLENWERKE A-G and BRUNO REWALD. Ger. 522,041, Dec. 6, 1927. Addn. to 516,187 (*C. A.* 25, 2019). As an emulsifying

agent for prepg the compn described in Ger. 516,187, use is made of residues from the extra of soy-bean phosphatides, *e g.*, of dregs from soy bean oil

Impregnating leather. *Le cuir lissé français*. Fr. 698,720, Oct. 11, 1929. Leather is impregnated with a neutral mixt. of paraffin, petrolatum and boiled or crude linseed oil in an non inflammable solvent, particularly CCl_4 .

Ornamenting artificial leather, oil cloth, wood, etc. ADOLF ESSLER. Austrian 121,989, Nov. 15, 1930. Addn. to 114,418 (C. A. 24, 987). The materials are treated as described in Austrian 114,418

Degradation products of albumins. CHEM. FAB. GERNAU LANDSHOFF & MEYER A-G (Linch Daring, inventor) Ger. 522,061, June 30, 1928. Leather waste, *e g.*, chrome leather shavings, is heated with an aq soln of an alk. earth hydroxide, and the mixt. is then treated with an alkali carbonate in an amt. insufficient to react with all the alk. earth hydroxide. Heating is continued, and the soln is then filtered, treated with Na_2CO_3 to ppt. CaCO_3 , and filtered again

Demineralizing bones. GELATINES HASSELT & VILORDE. Fr. 697,634, June 18, 1930. See Belg. 371,121 (C. A. 25, 1704).

30—RUBBER AND ALLIED SUBSTANCES

C. C. DAVIS

Evaluating gas blacks by the D. P. G. [diphenylguanidine] adsorption method. I. DAOEN. *India Rubber World* 83, No. 6, 57-9, 62 (1931) —Not all gas blacks adsorb a given accelerator to the same extent, and therefore the rate of vulcanization of a rubber mixt. contg. the accelerator varies with the gas black used. The analytical method described is claimed to be reliable in evaluating the effect of gas blacks on the rate of vulcanization and on the phys. properties of the vulcanizate, with diphenylguanidine as accelerator. *Procedure* —Agitate 2 g. of gas black with 100 cc. of MeOH solns of diphenylguanidine of 6 different normalities (0.0025-0.20 N), centrifuge and titrate 10 cc. of the supernatant liquid with MeOH solns of HCl of the corresponding normalities (bromophenol as indicator). All calcs. are expressed volumetrically in terms of a single normality of HCl. The quantity of diphenylguanidine adsorbed as a function of the quantity remaining in soln. (both in terms of HCl), when expressed in graphical form (ordinate and abscissa, resp.) gives a curve which, converted to a logarithmic basis, becomes a straight line. The distance between the abscissa axis and the point of intersection of this straight line with the ordinate axis represents the logarithm of a const. The antilogarithm of this const. represents the value of the gas black. Exptl. data show the application of this method to 4 different gas blacks. C. C. DAVIS

Behavior of various clays with crude and reclaimed rubber. H. A. WINKELMANN AND E. G. CROCKMAN. *Ind. Eng. Chem.* 22, 865-9 (1930) —Often the influence of 2 different clays on the properties of a rubber mixt. is greater in the uncured than in the cured mixt. The present paper shows the variations in raw rubber and in reclaimed rubber produced by various com. clays, as judged by plasticity, retentivity, phys. tests, artificial aging and chem. analysis. A microscopic examn. also shows the relation between particle size and phys. properties of the raw and reclaimed rubbers. The results are described in detail and shown graphically, and lead to certain general conclusions: (1) Chem. analysis of a clay gives little or no useful information about its behavior in raw or reclaimed rubber. On the other hand, its acidity or alkali has an influence on the rate of vulcanization and should be detd. The color of a clay does not indicate its purity or behavior. (2) There are wide differences in the effects of different clays on the plasticity, retentivity and softness of raw and reclaimed rubbers. (3) In cured new and reclaimed rubbers, clays vary greatly in their effect on the phys. properties, the same relative effects being obtained in the new and reclaimed rubber mixts. There is not, however, so much variation in tensile strength in reclaimed rubber vulcanizates as in new rubber vulcanizates. (4) Artificial aging in compressed O serves better to distinguish clays and is an important test in judging their merits. (5) By microscopic examn. it is possible to judge the relative effects of clays in new and reclaimed rubber, both uncured and cured. C. C. DAVIS

Rosin and rosin oil in rubber and reclaimed rubber. H. A. WINKELMANN AND E. B. BUSENBURG. *Proc. Am. Soc. Testing Materials* 30, Pt. 11, 807-27 (1930) —Rosin and rosin oil can be satisfactorily used as softeners for rubber, with due regard to proper compounding conditions. Different types and grades of rosin and rosin oil are described, and the phys. properties of both gum and wood resins are given in tables.

"Synthetic" rosin oils of low rosin content, about pine tar and kila pine tar, having acid nos. of 65-90, can be used interchangeably as softeners, but rosin oils containing much rosin, rosin itself or neutral rosin oil must each be handled differently. Besides acting as a softener during the mfg. process, rosin also activates the accelerators, but because of its relatively greater effect as a softener the cure is slower with rosin than with fatty acids. To prevent rosin from lowering the modulus and tensile strength of rubber, greater proportions of accelerator must be used. Three different grades of rosin appeared to affect modulus, tensile strength and aging equally. If wood rosin is substituted for the pine tar in an ordinary tread stock formula, the rate of vulcanization and tensile strength are slightly lowered, and if substituted for the stearic acid there is considerable softening and the rate of vulcanization is lowered. A no. of expts. are described, in which various grades of gum and wood rosin, pine tar and "synthetic" rosin oil are incorporated as softeners in a rubber compn. and their effect on plasticity, "nerve," elongation and tensile strength, aging and softness ascertained. Rosin is also useful in imparting tackiness to rubber, and as a devulcanizing agent. Also in *India Rubber World* 83, No. 6, 635, 67(1931), 84, No. 1, 568(1931). H. K. SALZBERG.

Some recent engineering applications of rubber. J. R. HOOVER AND F. L. HAUSHALTER. *Ind. Eng. Chem.* 23, 4629(1931).—The discussion includes rubber bearings, electrodeposited rubber products, absorption of vibration and noise, heat resisting and oil resisting rubber products, rubber products resistant to chemicals, rubber as a preventive of aviation ice hazard and various other applications. C. C. DAVIS.

Toxic substances in the rubber industry. XX. Sulfur monochloride. P. A. DAVIS. *Rubber Age* (N. Y.) 29, 778(1931), cf. C. A. 25, 2329.—A review and discussion of the physical properties of S_2Cl_2 and preventive and remedial measures to be taken. C. C. DAVIS.

Accelerated discoloration tests for white rubber products and methods of measuring the extent of the discoloration. HARLAN A. DREW. *Rubber Age* (N. Y.) 29, 27-30(1931).—The discoloration (yellowing) of vulcanized rubber upon exposure to sunlight is a function of (1) the character of the light (the shorter the wave length the more rapid the discoloration), (2) the temp. (the higher the temp. the greater the discoloration) and probably (3) moisture (the yellow color is leached away by water). Light from a Hg arc may be used as the basis of an artificial discoloration test, and the changes in color may be measured by a "K and R" color analyzer. In this way it was shown that the higher the lithopone content of vulcanized rubber the less did the latter discolor, and that ZnO inhibited the yellowing of vulcanizates containing lithopone. C. C. DAVIS.

The effect of grit upon the stress and strain properties of a carbon-black stock. I. P. W. KIRKLEY AND C. R. PARK. *Rubber Age* (N. Y.) 29, 79-80(1931).—From a practical point of view, a small proportion of grit in C black does not impair the resistance to abrasion of a vulcanizate containing it, as shown by Twiss (cf. C. A. 20, 3338), and it is only the ultimate elongation and tensile strength which may be seriously affected. The present paper describes expts. on this latter effect, a field of investigation not well covered in the literature. To the base mixt. rubber 100, S 35, diphenylguanidine 1.25, ZnO 5, were added samples of a C black to which had been added increasing proportions of C-black grit from elsewhere. The mixts. were cured 20, 60 and 90 min. at 140° and tested for tensile strength and modulus. Up to 3% the grit had no measurable effect on the tensile strength or on the modulus. Above this proportion, the values began to decrease progressively. When poor results are obtained with gas black, factors other than grit are responsible, notably poor dispersion of the pigment (cf. Endres, C. A. 19, 1208). Agglomerates in the unvulc. state disintegrate upon stretching and form points of weakness. The present dry tendency to minimize times of mixing results in incomplete dispersion, and it is probable that premature tread wear and cracking are a result of this rather than of gritty pigments. C. C. DAVIS.

A contribution to the problem of the impregnability of cord threads with rubber. F. A. HAUSSER AND M. HENNEBERGER. *Metallgesellschaft Periodic Ber.* No. 5, 138(1931).—There is almost no literature on the impregnability of fabrics by latex and by rubber in org. solvents except the investigations of Dieterich (C. A. 24, 1543) and Grenquist (C. A. 22, 4875), and the methods used by both give untrustworthy results because of the changes which may take place during the rigorous treatment. For this reason a new microscopic method was developed which should be free of these shortcomings and with which were examd. fabric or cord (1) covered with visibly thick rubber, e. g., coated on both sides with or without previous impregnation, and (2) protected by a rubber coating of no significant thickness, e. g., impregnated with latex or rubber in an org. solvent. In case (1) the sample is set in hard paraffin, dissected and

frozen with CO_2 to a depth of 20μ , a microsection $20-30\mu$ thick is cut on a sliding microtome, immersed in warm water, dried on a microscope slide, immersed in concd H_2SO_4 for 1-2 hrs., washed clean with H_2SO_4 and examd directly or preserved in Canada balsam. In case (2) the sample is immersed in 7% gelatin soln and solidified by cooling, and the same procedure used as before. The samples thus examd were: (1) thread soaked in 5% benzine rubber soln, (2) sample (1) after treatment with H_2SO_4 , (3) thread soaked in 33% latex (ammomated), (4) sample (1) after treatment with H_2SO_4 , (5) dry cord threads frictioned with rubber, (6) sample (5) after treatment with H_2SO_4 , (7) cord fabric soaked in rubber in $\text{C}_{11}\text{H}_{16}$, then flamecoated on both sides and finally treated with H_2SO_4 , (8) latex cord fabric coated on both sides and treated with H_2SO_4 , (9) cord fabric coated on both sides and after vulcanization treated with H_2SO_4 , (10) latex cord fabric coated on both sides and after vulcanization treated with H_2SO_4 , and (11) dry cord fabric in sheets treated with H_2SO_4 after vulcanization. Microscopic examn of these different samples made possible, in conjunction with earlier expts on phenomena taking place during vulcanization (cf *C A* 22, 4876, 23, 3597), the drawing of certain general conclusions. A fabric previously soaked in a $\text{C}_{11}\text{H}_{16}$ soln of rubber is already completely impregnated with rubber before vulcanization, whereas a fabric soaked in latex is only covered with an external coating of rubber, even after vulcanization. When threads which have not been treated are embedded in rubber sheets, or when rubber mixts are frictioned on dry fabrics, there is no impregnation previous to vulcanization, but impregnation takes place to a high degree during vulcanization as a result of the flow of the rubber. The extent to which wireless cords can be impregnated in a vulcanized state depends not upon whether the fabric has been previously treated but upon the phys condition of the rubber. The reason that the rubber in latex is merely deposited on the surface is that the rubber hydrocarbon is present as particles whose av diam is $1-2\mu$, and the capillary interstices of the fibers are smaller than these particles. The question whether cord fabrics and wireless cords prep'd with $\text{C}_{11}\text{H}_{16}$ solns of rubber or those with latex are superior in strength, elasticity and durability is not settled by the present investigation. C. C. DAVIS

Graphical tensile-testing machine for rubber threads. S. H. HAHN AND E. O. DIETERICH *Ind Eng Chem, Anal Ed* 3, 214-21 (1931).—The new machine, the construction and operation of which are described in detail, allows the testing of thread having breaking strengths of approx 400-2000 g. and elongations up to 1000%, with great precision at 200-700% elongation. Curves show the stress-strain curves of typical threads by this method and a comparison in 1 case with the stress-strain curve of the same vulcanizate by the ordinary method. C. C. DAVIS

Experiments on some technical mixtures with reference to their electrical properties. I. JACOBS *Kautschuk* 7, 4-7 (1931).—The object of the expts. was to ascertain to what extent the compn and the color of rubber matting influence the elec insulating power. In expts already described (cf *C A* 24, 4658) it was found that certain coloring agents like ochre and C black have unfavorable effects on the resistance to elec puncture. In places where the sheets were united during vulcanization, and at the junctions of 2 colors in mottled samples, the resistance to breakdown was no lower than in the corresponding homogeneous sections. The higher the voltage the shorter the time before breakdown. A systematic study of the influence of various fillers on the dielec resistance of matting was carried out by using a base mixt contg 45% natural whitening and comparing the behavior of this with corresponding mixts contg kaolin, kieselguhr, slate powd, marble and pptd CaCO_3 instead of whitening. With kaolin and with pptd CaCO_3 , the resistance to elec breakdown was less than with the other fillers. Kieselguhr and powd marble gave the highest resistances, followed by slate and natural whitening, which were greatly inferior to kieselguhr and powd marble. In every case the green matting samples contg CaO , were inferior to the gray ones. The quant results of the elec measurements are given in tables. C. C. DAVIS

Accelerators of vulcanization. I. JACOBS *Caoutchouc & gutta-percha* 25, 15438-45 (1931), cf *C A* 25, 2330.—The manuf, phys and chem properties and behavior from an accelerating point of view of com thuram derivs are described, with quant data. *Ibid* 15485-91.—The properties and behavior of tetramethylthuram disulfide under various conditions and in various types of rubber products are described, with quant data. C. C. DAVIS

An outline of the manufacture of carbon black (CRANOR, VENUTO) 18. Polymerizing oils and fats [for rubber industry] (Fr. pat 697,785) 27.

Rubber compositions. PAUL HOYER AND HANS HOYER *Ger.* 522,091, Aug 25,

1926 A fusible rubber compound that assumes its original condition on cooling is prepared by heating rubber or gutta-percha with about 20–25% of paraffin, adding S to the melt in the proportion required for vulcanization and further heating until a liquid of watery consistency is obtained, to which fillers, adjuvant rubber and S, etc., may be added. The mixt. should be stirred throughout the process. The product may be cast, or applied to containers as a protective lining. It contains only a little paraffin, the greater part evap. during the process. An example is given.

Sponge rubber. GUMMI UND BALATA WERKE "MATADOR" A. G. Austrian 122,000 Nov. 15, 1930. Variegated sponge rubber showing marbled or stained effects is prepared by combining a no. of differently colored rubber masses to each of which suitable softening and gas-forming reagents have been added. The masses may be drawn out into thin sheets and superposed, then rolled up together, twisted and rolled out into strips, which are lathe-sol. by sink in a mold and vulcanized.

Preventing the deterioration of rubber. THE NAUGATUCK CHEMICAL CO. Fr. 197,630 June 18, 1930. Rubber and other vulcanizable materials are treated with products obtained by the reaction in the presence of a dehydrating agent of a secondary amine having the general formula $R-NH-R'$ with an acid of the general formula $RCOOH$ in which R is an org. radical or H and R' and R are mono- or polycyclic residues in each of which the C atom α to the N atom is joined to a H atom. Methylacridium obtained by the reaction of $PhNH_2$ on $Al(OH)_3$ naphthalenacridines from phenyl-naphthylamine or diacetylstyrene may be used.

Inhibiting cracking of rubber when exposed to sunlight. SIDNEY M. CADWILL and LEONARD MAYER (to Naugatuck Chemical Co.) U. S. 1,798,131 March 31, 1930. Sulfonates (suitably 1–5% the quantity of the rubber) are used such as diphenyl sulf oxide, di-p-tolyl sulf oxide, dibenzyl sulf oxide, dimethyl sulf oxide or dibenzyl disulf oxide.

Antioxidant for treatment of rubber. SIDNEY M. CADWILL and BIERMAN STRICKHOUSER (to The Dominion Rubber Co. Ltd.) Can. 310,115, Apr. 7, 1931. Deterioration and surface cracking of rubber are retarded by treating vulcanizable rubber stock with the material obtained by mixing a poly(ethylene)pyrrolone b. above 130° and a monohydroxynaphthalene, and then vulcanizing the stock.

Antioxidant for treatment of rubber. CLYDE COLEMAN (to The Dominion Rubber Co., Ltd.) Can. 310,141, Apr. 27, 1931. The deterioration of rubber is retarded by treating rubber in the presence of an accelerator and a vulcanizing agent with 1,4'-diaminodiphenylmethane, and subsequently vulcanizing the rubber.

Regenerated rubber. AMERICAN GLUE CO. Fr. 698,452, July 7, 1930. Rubber is recovered from waste rubber by submitting the waste to the action of the cleavage products from the disson of tanned leather during the process of devulcanization.

Plastic materials. JEAN BARR. Fr. 697,611, June 18, 1930. A plastic-elastic material is made by dissolving in CS_2 the polymerization products of halogenated hydrocarbons of the group $C_{12}H_{10}$, with polysulfides and by mixing this soln. with natural rubber latex. Cf. C. A. 24, 628.

Masses resembling hard rubber. I. G. FARRENHED A.-G. Fr. 698,208, July 2, 1930. Mixts. of natural rubber and artificial masses resembling rubber are treated in the same way as hardened rubber is usually treated. Besides S vulcanization accelerators and fillers may be added.

Temperature-regulating system for rubber extrusion apparatus. RALPH B. DAY (to Goodyear Tire & Rubber Co.) U. S. 1,800,180, April 7.

Rubber coatings. DUNLOP RUBBER CO., LTD., and THE ANODER RUBBER CO., LTD. Fr. 697,762, June 21, 1930. Cellulosic or fibrous vessels are coated with rubber or similar vegetable resins applied in the form of solns. or aq. emulsions or dispersions, which may also contain a wax or mixt. of waxes.

Hose of rubber and fibrous material. CHARLES W. LEQUILLON (to B. F. Goodrich Co.) U. S. 1,798,798, March 31. Various details of mech. assembly and vulcanization are described.

Compounding soap-forming materials with other substances. WILLIS A. GIBBONS (to American Rubber Co.) U. S. 1,798,253, March 31. In forming products suitable for compounding with rubber latex, a water-sol. volatile base soap of a higher aliphatic acid such as NH_4 stearate is first mixed with a filler such as C black, in finely divided fluid form, which is nonreactive with the acid, the material is mixed with a rubber latex, and the soap is then reconverted into the original material (suitably by removing NH_4 in a drying operation).

Synthetic rubber. I. G. FARRENHED A.-G. Ger. 522,090, July 13, 1928. Butadiene hydrocarbons are polymerized by treatment with NaH at atm. or raised temp. An example is given.

Synthetic rubber. I G FARBERMAN A-G Fr 697,679, June 19, 1930 In the polymerization of 1,3 butadiene or its homologs or analogs by the aid of alkali metals, the metal is placed in a vessel with one or more walls and is in communication through openings of at least 3 mm diam with the hydrocarbon. The vessel used may be filled up with the hydrocarbon which is allowed to polymerize and then added to the main mass of hydrocarbon.

Rubber-vulcanization accelerator. RALPH V. HAUSKE (to American Cyanamid Co.) U S 1,798,159, March 31. Dibenzyldiamine is used as an accelerator and general mention is made of the use of some similar compounds.

Rubber-vulcanization accelerators. THE NAUATUCK CHEMICAL CO. Fr 698,259, July 1, 1930 The aliphatic bases produced by the reaction of dialkyls of ethylene or its homologs with NH_3 are used as accelerators.

Triazine derivatives containing sulfur. SOC ANON, POUR L'INDUSTRIE A BALE. Fr 697,599, June 18, 1930 Triazine derivs contg at least one atom of exchangeable halogen are condensed with at least one mol of a S deriv. contg an SH group and a group easily removed. Such derivs of S are thiosulfates, thiocyanates, xanthogenates, di or poly sulfides. The products are transformed by sapon and reduction to mercapto triazines. Examples are given including the prepn of dimercaptophenylamino triazine (m 248°), mercaptoaminophenylamino triazine (m 240°), dimercaptanaphthylamino triazine (m 269-262°) and mercaptoamino- α -naphthylamino triazine (m. 257-258°). The products are used in the vulcanization of rubber.

Apparatus for vulcanizing tire and other rubber articles. PAUL WIRCHARDT (to Fr Krupp Grusonwerk A G) U S 1,798,826, March 31 Mech. features.

Vulcanizing rubber. I G FARBERMAN A G Fr 697,568 June 17, 1930 Mixts. of rubbers obtained in different ways from natural and synthetic rubbers are vulcanized in the presence of finely divided C (lampblack) Cf C. A. 25, 1412

CHEMICAL RESEARCH PERIODICALS

COMPLETE SETS. VOLUMES OR SINGLE
COPIES OF DOMESTIC AND FOREIGN
CHEMICAL & SCIENTIFIC MAGAZINES.
BOUGHT & SOLD.

B. LOGIN & SON, INC.

Estab. 1897

29 East 21st St., New York

Professional Directory Announcements in

*Industrial and
Engineering Chemistry*
constitute
dignified publicity.

One inch card \$4.40
per insertion,
\$48.00 a year.

Two inch card \$8.80
per insertion,
\$96.00 a year.

CHEMICAL ABSTRACTS and other TECHNICAL PERIODICALS

Scientific Journals, reports,
proceedings, etc., published in
United States or Europe,
bought and sold. Complete
sets, volumes or single copies.
Catalog on request.

THE H. W. WILSON CO.
958-964 University Ave.
New York, N. Y.
U. S. A.